

Lewis Acid Reagents

A Practical Approach

Edited by
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OXFORD
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Published in the United States
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ISBN 0-19-850099-8

Printed in Great Britain by

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Preface

Lewis acids are becoming a powerful tool in many different modern reactions, such as the Diels-Alder reactions, Ene reactions, Sakurai reactions, and Aldol synthesis. In fact, the importance and practicality of Lewis acid reagents as valuable means of obtaining a variety of organic molecules is now fully acknowledged by chemists in the synthetic organic society. This prominence is due to the explosive development of newer and even more efficient methods during the last decade, and the numbered publications on these reagents is actually increasing exponentially each year. Research on asymmetric synthesis has become more important and popular in the total synthesis of natural products, pharmaceuticals, and agricultural agents, and Lewis acid chemistry plays a major role in this arena.

Comprehensive coverage of the literature on each area of Lewis acid is not necessarily provided here. Rather, the aim of the book is to furnish a detailed and accessible laboratory guide useful for researchers who are not familiar with the benefits of Lewis acids. It includes information on reagent purification, reaction equipment and conditions, work-up procedures, and other expert advice. The primary goal is thus to dispel the mystery surrounding Lewis acid reagents and to encourage more scientists to use these powerful synthetic tools to maximum effect. The book contains 14 independently referenced chapters describing a variety of Lewis acids using different metals. Each metal has different characteristic features of reagent preparation and practicality which clearly described in that chapter.

I would like to thank Professor K. Ishihara for helping to check parts of the manuscripts and for useful suggestions. I would also like to express my personal gratitude to all of the invited contributors who carefully honored the deadlines and thus made the editorial job much easier.

It is my strong hope that this book will be found an invaluable reference for graduate students as well as chemists at all levels in both academic and industrial laboratories.

Nagoya
October 1998

H. Y.

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Contents

| | |
|--|------|
| <i>List of contributors</i> | xiii |
| 1. Introduction | 1 |
| <i>Hisashi Yamamoto</i> | |
| 2. Synthetic utility of bulky aluminium reagents as Lewis acid receptors | 5 |
| <i>Keiji Maruoka</i> | |
| 1. Introduction | 5 |
| 2. Amphiphilic alkylations | 6 |
| 2.1 Amphiphilic carbonyl alkylations | 6 |
| 2.2 Amphiphilic conjugate alkylations | 10 |
| 2.3 Amphiphilic conjugate allylations | 12 |
| 3. Regio- and stereocontrolled Diels–Alder reaction | 15 |
| 4. Asymmetric hetero-Diels–Alder reaction | 16 |
| 5. Claisen rearrangement | 21 |
| 6. Epoxide rearrangement | 23 |
| References | 28 |
| 3. Boron reagents | 31 |
| <i>Kazuaki Ishihara</i> | |
| 1. Introduction | 31 |
| 2. Arylboron reagents with electron-withdrawing aromatic group | 32 |
| 2.1 Tris(pentafluorophenyl)boron as an efficient, air stable, and water-tolerant Lewis acid catalyst | 32 |
| 2.2 Diarylborinic acids as efficient catalysts for selective dehydration of aldols | 35 |
| 2.3 3,4,5-Trifluorophenylboronic acid as an amidation catalyst | 37 |
| 3. Chiral boron reagents as Lewis acids | 39 |
| 3.1 Catalytic enantioselective carbo-Diels–Alder reactions | 39 |

| | |
|--|-----------|
| 3.2 Catalytic enantioselective aldol and allylation reactions | 51 |
| 3.3 Asymmetric reactions of imines mediated by chiral boron reagents | 57 |
| References | 61 |
| 4. Magnesium(II) and other alkali and alkaline earth metals | 65 |
| <i>Akira Yanagisawa</i> | |
| 1. Introduction | 65 |
| 2. Cycloaddition reactions | 65 |
| 3. Aldol additions, allylations, and other reactions | 71 |
| References | 74 |
| 5. Zinc(II) reagents | 77 |
| <i>Nobuki Oguni</i> | |
| 1. Introduction | 77 |
| 2. Scope and limitations | 77 |
| References | 88 |
| 6. Chiral titanium complexes for enantioselective catalysis | 93 |
| <i>Koichi Mikami and Masahiro Terada</i> | |
| 1. Introduction | 93 |
| 2. Carbonyl addition reaction | 96 |
| 3. Carbonyl-ene reaction | 101 |
| 4. Asymmetric catalytic desymmetrization | 105 |
| 5. Kinetic optical resolution | 106 |
| 6. Positive non-linear effect of non-racemic catalysts | 107 |
| 7. Enantiomer-selective activation of racemic catalysts | 110 |
| 8. Ene cyclization | 111 |
| 9. Aldol reaction | 113 |
| 10. (Hetero) Diels–Alder reaction | 118 |
| 11. Cyanohydrine formation | 129 |
| 12. Miscellaneous reactions | 130 |
| Acknowledgements | 131 |
| References | 131 |

| | |
|---|-----|
| 7. Tin Lewis acid | 137 |
| <i>Shū Kobayashi</i> | |
| 1. Introduction | 137 |
| 2. Asymmetric aldol reaction | 138 |
| 2.1 Preparation of chiral tin(II) Lewis acids | 138 |
| 2.2 Asymmetric aldol reactions of silyl enol ethers derived from acetic and propionic acid derivatives | 138 |
| 2.3 Asymmetric synthesis of 2-substituted malate derivatives | 140 |
| 2.4 Asymmetric synthesis of <i>syn</i> - and <i>anti</i> -1,2 diol derivatives | 143 |
| 2.5 Catalytic asymmetric aldol reactions | 147 |
| 3. Asymmetric cyanation reaction | 151 |
| 4. Asymmetric protonation | 153 |
| References | 155 |
| 8. Silicon(IV) reagents | 159 |
| <i>Akira Hosomi and Katsukiyo Miura</i> | |
| 1. Introduction | 159 |
| 2. Carbon–oxygen and carbon–nitrogen bond formation | 160 |
| 3. Carbon–carbon bond formation | 161 |
| 4. Carbon–hydrogen bond formation | 164 |
| References | 167 |
| 9. Silver and gold reagents | 169 |
| <i>Masaya Sawamura</i> | |
| 1. Introduction | 169 |
| 2. Enantioselective aldol reactions of activated isocyanides and aldehydes | 169 |
| 3. Enantioselective aldehyde allylation with allyltin reagents catalysed by a silver(I)–phosphine complex | 173 |
| References | 175 |
| 10. Zr- and Hf-centred Lewis acid in organic synthesis | 177 |
| <i>Keisuke Suzuki</i> | |
| 1. Introduction | 177 |

| | |
|---|------------|
| 2. Activation of C–F bond | 177 |
| 3. Activation of ethers | 179 |
| 4. Carbonyl activation | 180 |
| 5. Reductive coupling of allenes and alkynes | 182 |
| References | 184 |
| 11. Scandium(III) and yttrium(III) | 185 |
| <i>Shin-ichi Fukuzawa</i> | |
| 1. Introduction | 185 |
| 2. Scandium(III) triflate | 185 |
| 2.1 Friedel–Crafts reaction | 185 |
| 2.2 Acetalization | 189 |
| 2.3 Acetylation and esterification of alcohol | 190 |
| 2.4 Allylation of carbonyl compounds with tetra-allyltin catalysed by scandium(III) triflate | 192 |
| 2.5 Scandium(III) triflate-catalysed Mukaiyama aldol reaction | 194 |
| 2.6 Diels–Alder reaction | 196 |
| 3. Cyclopentadienyl yttrium hydride | 199 |
| References | 200 |
| 12. Lanthanide(III) reagents | 204 |
| <i>Takeshi Nakai and Katsuhiko Tomooka</i> | |
| 1. Introduction | 204 |
| 2. Cerium(III) reagents | 204 |
| 3. Europium(III) reagents | 208 |
| 4. Ytterbium(III) reagents | 214 |
| 5. Lanthanum(III) reagents | 217 |
| 6. Other lanthanide(III) reagents | 219 |
| References | 221 |
| 13. Other transition-metal reagents: chiral transition-metal Lewis acid catalysis for asymmetric organic synthesis | 225 |
| <i>Hisao Nishiyama and Yukihiro Motoyama</i> | |
| 1. Introduction | 225 |

Contents

| | |
|---|-----|
| 2. Activation of haloalkanes and carbonyl compounds | 225 |
| 3. Catalytic asymmetric Diels–Alder reactions | 228 |
| 3.1 Achiral transition-metal catalysts | 228 |
| 3.2 Chiral transition-metal catalysts | 229 |
| 4. Asymmetric aldol condensations | 238 |
| References | 242 |
| | |
| 14. Lewis acid-assisted anionic polymerizations for synthesis of polymers with controlled molecular weights | 246 |
| <i>Takuzo Aida and Daisuke Takeuchi</i> | |
| 1. Introduction | 246 |
| 2. Principle, scope, and limitations of ‘Lewis acid-assisted high-speed living anionic polymerization’ | 247 |
| References | 260 |
| | |
| Index | 263 |

Introduction

HISASHI YAMAMOTO

Are enzymatic reactions really good models for laboratory chemical reactions? An enzyme is a giant molecule, large enough to support a substrate, whereas chemical reagents are composed of much smaller molecules. Still, the much smaller molecular apparatus of human-made reagents is expected to induce reactions with selectivities comparable to those of a large enzyme. Clearly, the design of new reagents requires careful abstraction and simplification of the true mechanism of an enzyme, much like the design of an aircraft might be based on the aerodynamics of a bird.

A case in point is the important role of hydrogen bonding during enzymatic reactions. In the course of such processes, the giant template of the enzyme will specify quite accurately the position and direction of a proton for hydrogen bonding, before and after the reaction. However, a proton by itself cannot behave in this fashion. A perfect sphere, it has no directional selectivity for hydrogen bonding outside the domain of the enzyme, thus it is unable to act as a 'delicate finger' in an ordinary organic reaction as it does in the enzymatic transformation. It is natural to wonder whether an appropriate substitute for the proton might induce human-made reactions capable of selectivities comparable to those afforded by enzymes.

An excellent candidate as a proton substitute is a Lewis acid. The observation that organoaluminium, organolithium, organoboron and many other organometallic compounds immediately ignite when exposed to air, reflecting the high affinity of these metals for oxygen, inspired us to devise a new series of reagents based on those metals: true 'Designer Lewis Acids' for organic synthesis. For example, since an organoaluminium compound would have three ligands around the metal, the structural design of such a catalyst could be quite flexible. The goal, then, was to engineer an artificial proton of a special shape, which could be utilized as an effective tool for chemical reactions, by harnessing the high reactivity of the metal atom towards oxygen. Such a concept was initially researched by examining the influence of specially designed Lewis acid compounds.

In the *Encyclopedia of reagents for organic synthesis* edited by Paquette,

the reagent function index listed the following metals as being used as Lewis acid reagents:¹

Aluminum, Antimony, Boron, Cadmium, Cerium, Cobalt, Copper, Europium, Germanium, Hafnium, Iron, Lanthanum, Lithium, Magnesium, Molybdenum, Nickel, Palladium, Phosphorus, Silicon, Silver, Sulfur, Thallium, Tin, Titanium, Vanadium, Ytterbium, Zinc, Zirconium

A truly varied group of elements are used as the Lewis acid reagent and each metal has its own characteristic features. We therefore decided in this book to classify these reagents according to their metals.

It need not be pointed out that Lewis acid-promoted carbon-carbon bond formation reaction is one of the most important processes in organic synthesis. Classically, Friedel-Crafts reaction, ene reaction, Diels-Alder reaction, and Mukaiyama aldol synthesis are catalysed with ordinary Lewis acids such as AlCl_3 , TiCl_4 , $\text{BF}_3 \cdot \text{OEt}_2$, or SnCl_4 (Fig. 1.1). These classical Lewis acids activate the functional groups of substrates, and the reactions proceed in relatively low stereo-, regio-, or chemoselectivities. On co-ordination with well-designed ligand(s), a Lewis acid exhibits substantially new reactivity. Furthermore, a designer Lewis acid leads to an isolation of monomeric Lewis acid species whose structural features can be easily understood and easily extended to designer chiral catalysts for asymmetric syntheses. Thus, metal ligand tunings are the most essential component in the design of Lewis acid reagents.²⁻⁴

Lewis acid-mediated reactions can be classified as follows. The complex between substrate and Lewis acid rearranges to produced the product (type 1).

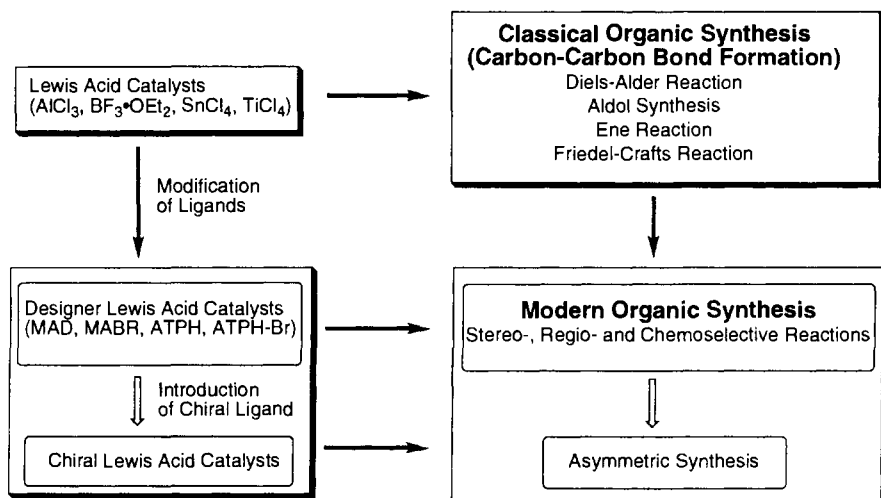


Fig. 1.1 Role of Lewis acid in organic synthesis.

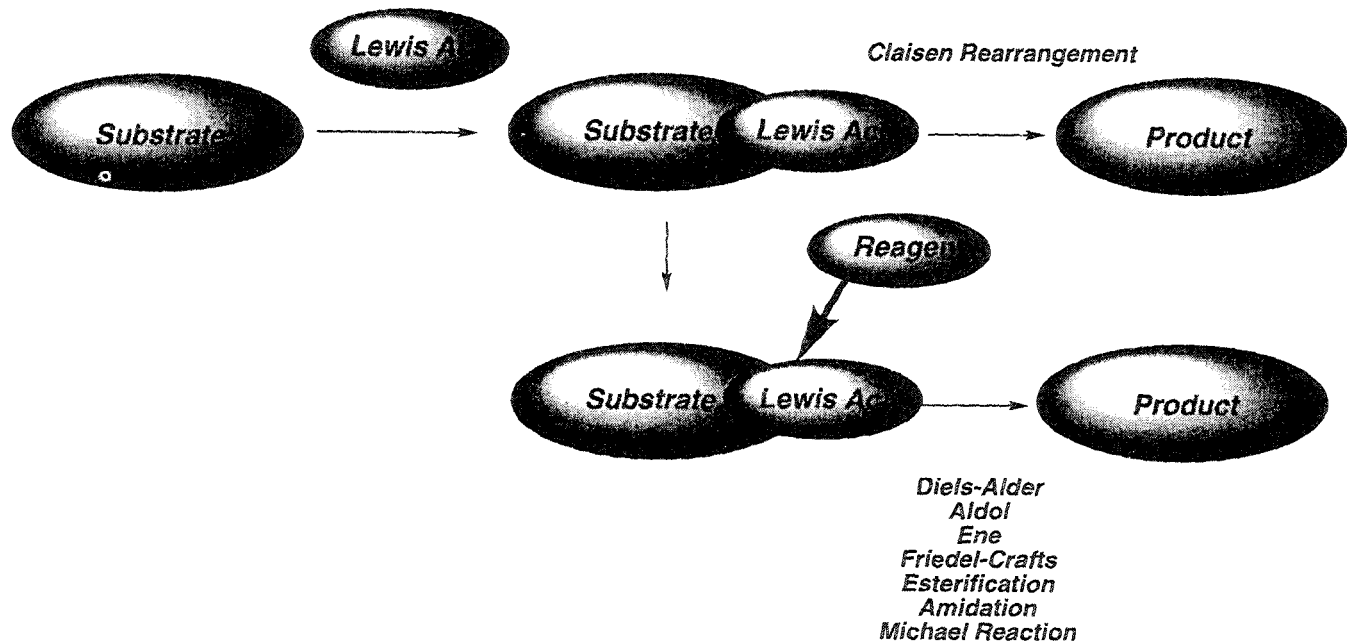


Fig. 1.2 Classification of Lewis acid promoted reactions.

Claisen rearrangement promoted by Lewis acid catalyst is a typical example of type 1. On the other hand, some complexes between Lewis acids and substrates are stable enough and the formed complexes react with a variety of reagents from outside the system to generate the product (type 2). The reaction between the Lewis acid activated unsaturated carbonyl compounds with dienes, Diels–Alder reaction, is an example of type 2 (see Fig. 1.2).

A Lewis basic carbonyl group can be activated through co-ordination with a metal-centred Lewis acid, with profound reactivity and stereochemical consequences. In the context of asymmetric synthesis, many of the Lewis acid-mediated reactions are known to proceed with improved stereoselectivities as compared to their non-catalysed counterparts; very recently, a number of chiral Lewis acids have been used as remarkably efficient catalysts for carbonyl addition processes. Although the origins of many of the effects brought about by Lewis acids are still poorly understood, it is clear that the conformational preferences of the Lewis acid carbonyl complex are ultimately responsible for determining the stereochemical course of Lewis acid-mediated reactions.^{5,6}

References

1. Paquette, L. A. (ed.) *Encyclopedia of reagents for organic synthesis*, John Wiley & Sons, Chichester, New York, Brisbane, Toronto, Singapore, **1995**.
2. Schinzer, D. (ed.) *Selectivities in Lewis acid promoted reaction*, Kluwer Academic Publishers, Dordrecht, Boston, London, **1988**.
3. Sasntelli, M.; Pons, J.-M., *Lewis acids and selectivity in organic synthesis*, CRC Press, Boca Raton, New York, London, Tokyo, **1996**.
4. *Methods of Organic Chemistry* (Houben-Weyl), Additional and Suppl. Vol. of the 4th Edn, Vol. E 21b; *Stereoselective Synthesis* (ed. G. Helmchen, R. W. Hoffmann, J. Mulzer, E. Schaumann), Thieme, Stuttgart, **1995**.
5. Shambayati, S.; Crowe, W. E.; Schreiber, S. L. *Angew. Chem., Int. Ed. Engl.* **1990**, 29, 256.
6. Denmark, S. E.; Almstead, N. G. *J. Am. Chem. Soc.*, **1993**, **115**, 3133.

Synthetic utility of bulky aluminium reagents as Lewis acid receptors

KEIJI MARUOKA

1. Introduction

Organoaluminium compounds, little known until the 1950s, have been widely accepted and increasingly important in the field of industry and in the laboratory,¹⁻¹³ particularly after K. Ziegler and colleagues discovered the direct synthesis of trialkylaluminiums and their brilliant application to the polymerization of olefins.^{14,15} The chemistry of organoaluminium compounds has been understood in terms of the Lewis acidity of their monomeric species, which is directly related to the tendency of the aluminium atom to complete electron octets. Organoaluminium compounds possess a strong affinity for various heteroatoms in organic molecules, particularly oxygen. In fact, bond strength of aluminium and electronegative atoms such as oxygen is extremely strong; the bond energy of the Al-O bond is estimated to be 138 kcal mol⁻¹. In view of this high bond strength, most organoaluminium compounds are particularly reactive with oxygen and often ignite spontaneously in air. Accordingly, they easily generate 1:1 co-ordination complexes even with neutral bases such as ethers, which is in marked contrast with lithium and magnesium derivatives. Utilization of this property, commonly identified with 'oxygenophilicity', in organic synthesis allows facile reactions with hetero atoms particularly oxygen- and carbonyl containing compounds.

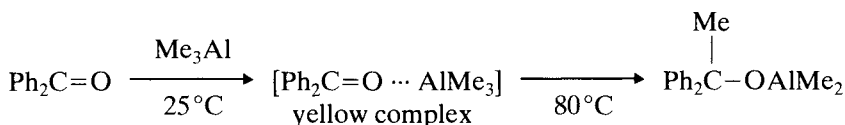


The major difference between organoaluminium compounds and more common Lewis acids such as aluminium chloride and bromide is attributable to the structural flexibility of organoaluminium reagents. Thus, the structure of an aluminium reagent is easily modified by changing one or two of its ligands. Described below are our recent practical strategies to selective organic synthesis with modified organoaluminium reagents.

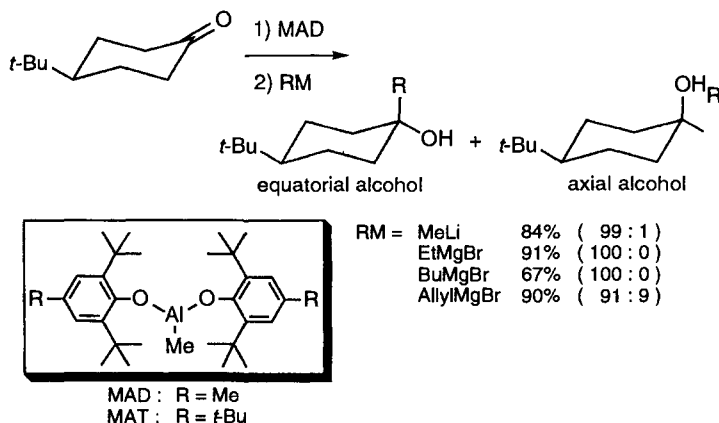
2. Amphiphilic alkylations

2.1 Amphiphilic carbonyl alkylations

Organoaluminium compounds are endowed with high oxygenophilic character, and hence are capable of forming long-lived monomeric 1:1 complexes with carbonyl substrates. For example, the reaction of benzophenone with Me_3Al in a 1:1 molar ratio gives a yellow, long-lived monomeric 1:1 species



which decomposed unimolecularly to dimethylaluminum 1,1-diphenylethoxide during some minutes at 80°C or many hours at 25°C .¹⁶ This unique property may be utilized for stereoselective activation of the carbonyl group. Among various organoaluminium derivatives examined, exceptionally bulky, oxygenophilic organoaluminium reagents such as methylaluminum bis(2,6-di-*tert*-butyl-4-alkylphenoxide) (MAD and MAT), have shown excellent diastereofacial selectivity in carbonyl alkylation.^{17,18} Thus, treatment of 4-*tert*-butyl cyclohexanone with MAD or MAT in toluene produced a 1:1 co-ordination complex which on subsequent treatment with methyl-lithium or Grignard reagents in ether at -78°C afforded the equatorial alcohol almost exclusively (Scheme 2.1). Methyl-lithium or Grignard reagents solely undergo preferential equatorial attack yielding axial alcohols as the major product. MAD and MAT have played a crucial role in the stereoselective synthesis of hitherto inaccessible equatorial alcohols from cyclohexanones as shown in Table 2.1.

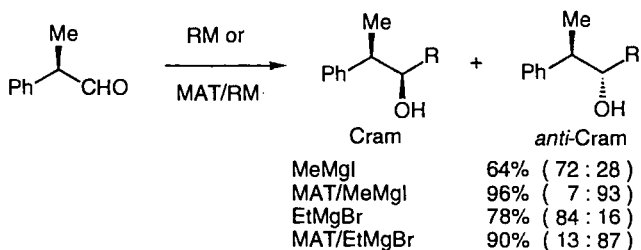
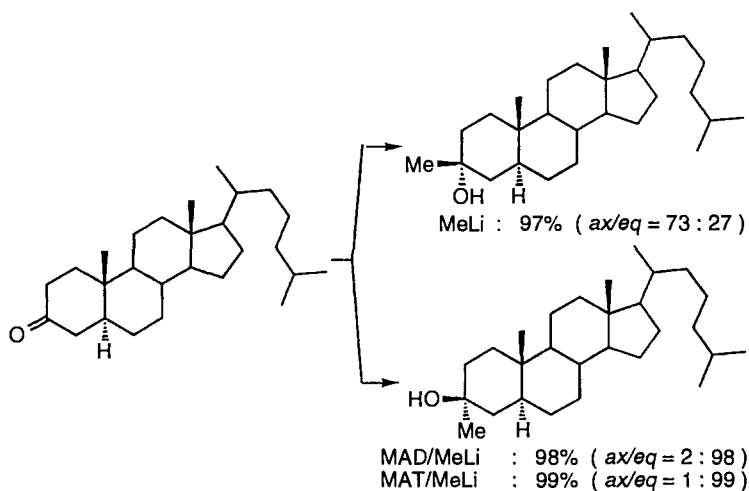


Scheme 2.1

Table 2.1 Stereoselective alkylation of cyclic ketones

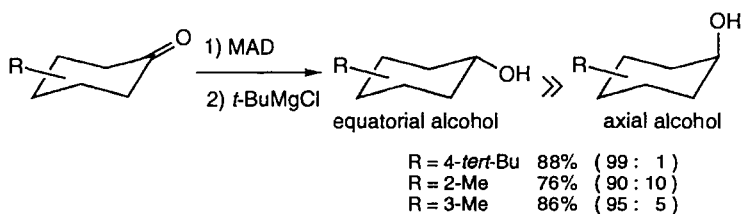
| Entry | Alkylation agent | Chemical yield (ax/eq ratio) |
|-------|------------------|------------------------------|
| | | |
| 1 | MeLi | 75% (79 : 21) |
| 2 | MAD/MeLi | 84% (1 : 99) |
| 3 | MAT/MeLi | 92% (0.5 : 99.5) |
| 4 | EtMgBr | 95% (48 : 52) |
| 5 | MAD/EtMgBr | 91% (0 : 100) |
| 6 | BuMgBr | 58% (56 : 44) |
| 7 | MAD/BuMgBr | 67% (9 : 100) |
| | | |
| 8 | MeLi | 73% (92 : 8) |
| 9 | MAD/MeLi | 84% (14 : 86) |
| 10 | MAT/MeLi | 80% (10 : 90) |
| | | |
| 11 | MeLi | 80% (83 : 17) |
| 12 | MAD/MeLi | 69% (9 : 91) |
| 13 | MAT/MeLi | 95% (3 : 97) |
| 14 | BuMgBr | 86% (79 : 21) |
| 15 | MAD/BuMgBr | 75% (1 : 99) |
| | | |
| 16 | MeLi | 77% (75 : 25) |
| 17 | MAD/MeLi | 82% (1 : 99) |

This approach has been quite useful in the stereoselective alkylation of steroidal ketones. Reaction of 3-cholestanone with MeLi gave predominantly 3 β -methylcholestan-3 α -ol (axial alcohol), whereas amphiphilic alkylation of the ketone with MAD/MeLi or MAT/MeLi afforded 3 α -methylcholestan-3 β -ol (equatorial alcohol) exclusively (Scheme 2.2). In addition, unprecedented *anti*-Cram selectivity was achievable in the MAD- or MAT-mediated alkylation of α -chiral aldehydes possessing no ability to be chelated.



Scheme 2.2

In contrast to the facile MAD- or MAT-mediated alkylation of cyclic ketones with primary organolithium or Grignard reagents, reduction takes precedence over alkylation with hindered alkylation agents such as *t*-butylmagnesium chloride in the presence of MAD¹⁸ (Scheme 2.3). This amphiphilic reduction system appears to be complementary to the existing methodologies using L-Selectride for obtaining axial selectivity.



Scheme 2.3

Protocol 1.

Synthesis of equatorial 4-*tert*-butyl-1-methylcyclohexanol. Amphiphilic alkylation of 4-*tert*-butylcyclohexanone with MAD/MeLi system

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

Equipment

- Magnetic stirrer
- Three-necked, round-bottomed flask (300 mL)
A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Teflon-coated magnetic stirring bar
- Medium-gauge needle
- All-glass syringe with a needle-lock Luer (volume appropriate for quantity of solution to be transferred)
- Vacuum/inert gas source (argon source may be an argon balloon)

Materials

- | | |
|--|--------------------------------|
| • 2,6-Di- <i>tert</i> -butyl-4-methylphenol (FW 220.4), 6.61 g, 30 mmol | irritant |
| • Trimethylaluminum in hexane (FW 72.1), 2 M solution in hexane, 7.5 mL, 15 mmol | Pyrophoric, moisture sensitive |
| • Dry toluene, 110 mL | flammable, toxic |
| • MeLi in ether (FW 22.0), 1.5 M solution in ether, 10 mL, 15 mmol | flammable, moisture sensitive |
| • 4- <i>tert</i> -butylcyclohexanone (FW 154.3), 1.54 g, 10 mmol | |
| • 1M HCl, 50 mL | toxic |
| • Technical ether for extraction, 100 mL | flammable, toxic |
| • Silica gel for flash chromatography 300 g, Merck Kieselgel 60 (Art. 9385) | irritant dust |
| • Ether for flash chromatography | flammable, toxic |
| • <i>n</i> -Hexane for flash chromatography | flammable, irritant |

1. Clean all glass wares, syringes, needles, and stirring bar and dry for at least 2 h in a 100°C electric oven before use.
2. Assemble the flask, stirring bar, and stop cocks under argon while the apparatus is still hot.
3. Support the assembled flask using a clamp and a stand with a heavy base.
4. Dry the apparatus with an electric heat gun under vacuum (1–2 mm Hg) for 5 min, then back-fill the flask with argon. Repeat to a total of three times.
5. Place 2,6-di-*tert*-butyl-4-methylphenol and flush with argon.
6. Charge dry toluene (100 mL).
7. Stir the mixture, degassed under vacuum, and replaced by argon.
8. Support the bottle containing trimethylaluminum in hexane using a clamp and a stand with a heavy base.
9. Fill a syringe with trimethylaluminum in hexane from the bottle containing trimethylaluminum using argon pressure. Apply the argon pressure to fill the syringe slowly with the required volume. Transfer the reagent in the syringe to the reaction flask at room temperature.
10. Stir the resulting solution at this temperature for 1 h to give methyl-

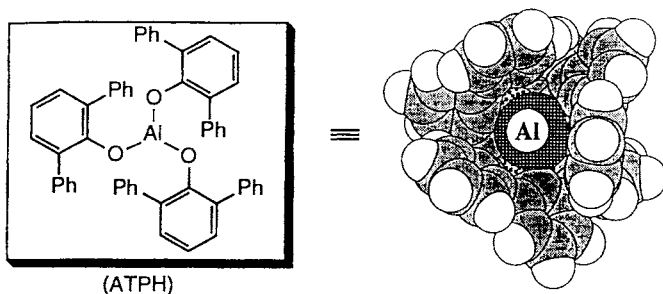
Protocol 1. Continued

aluminium bis(2,6-di-*tert*-butyl-4-methylphenoxide) (MAD) almost quantitatively. During this operation, nearly 2 equivalents of methane gas are evolved per 1 equiv of trimethylaluminium.

11. Cool the reaction vessel to -78°C in a dry ice-methanol bath.
12. Transfer 4-*tert*-butylcyclohexanone, which is dissolved in 10 mL of dry toluene, to the syringe, and add to the reaction flask over 5 min at -78°C . Then add MeLi in ether over 10–15 min at -78°C . Stir the whole mixture at -78°C for 1 h in order to complete the alkylation.
13. Place 1M HCl solution and the stirring bar in the Erlenmeyer flask. While 1M HCl solution is stirred vigorously, add the reaction mixture slowly at 0°C to avoid excessive foaming on hydrolysis.
14. Remove the ice bath and stir the entire mixture vigorously at 25°C for 20 min.
15. Transfer the mixture to a separating funnel and separate the two layers. Extract the water layer with ether twice (2×50 mL).
16. Combine the ethereal extracts in a 500 mL flask. Dry over anhydrous magnesium sulfate, and filter through filter paper. Concentrate the filtrate under reduced pressure by means of a rotary evaporator.
17. Purify the oily residue by column chromatography on silica gel (ether/hexane as eluants) to give 625 mg (84%) of a mixture of axial and equatorial 4-*tert*-butylcyclohexanols (ratio = 1:99) as a colourless oil, the ratio of which is determined by capillary GLC analysis. Each isomer is characterized by ^1H NMR analysis.

2.2 Amphiphilic conjugate alkylations

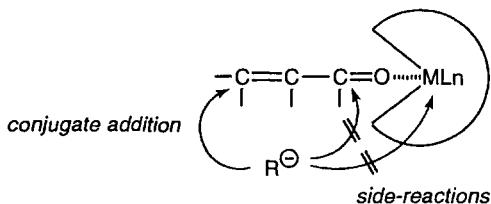
Conjugate addition to α,β -unsaturated carbonyl compounds is generally effected by soft organometallics as often seen in organocopper chemistry. In contrast, the use of organolithiums alone has never been developed to a useful level because of their hard nucleophilic character. This difficulty has been successfully overcome by using aluminium tris(2,6-diphenylphenoxide) (ATPH) (Scheme 2.4) as a carbonyl stabilizer.¹⁹ This type of conjugate alkyla-



Scheme 2.4

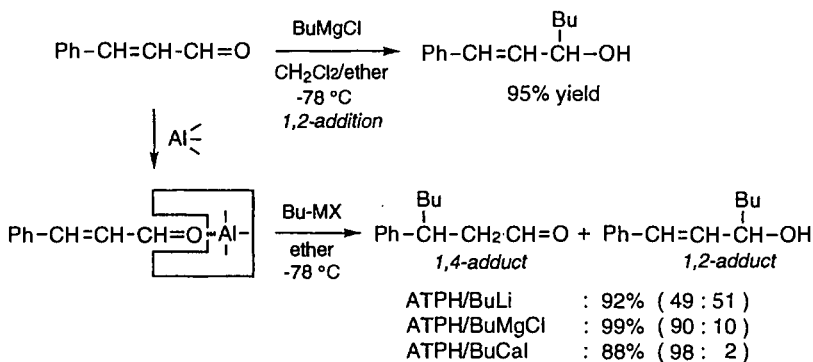
2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

tion can be classified as an amphiphilic conjugate alkylation that is markedly different from previously known nucleophilic and/or electrophilic conjugate alkylations (Scheme 2.5).

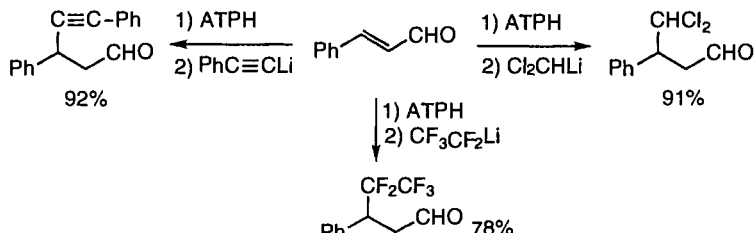


Scheme 2.5. Amphiphilic alkylation system.

This methodology is particularly effective for the conjugate alkylation to α,β -unsaturated aldehydes, which, among various conjugate acceptors, are prone to be more susceptible to 1,2 addition with a number of nucleophiles than α,β -unsaturated ketones, esters, and amides (Scheme 2.6). In addition, conjugate addition of lithium alkynides and thermally unstable lithium carbenoids, which are very difficult to achieve in organocopper chemistry, are realized with this amphiphilic conjugate alkylation system (Scheme 2.7).

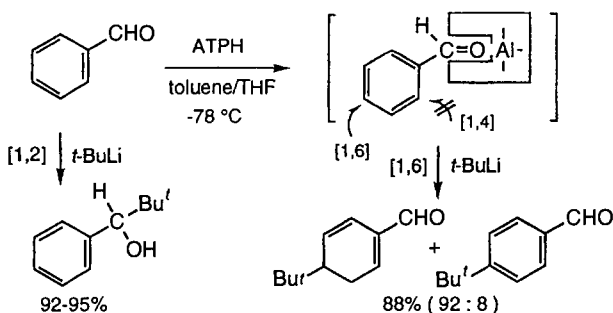


Scheme 2.6



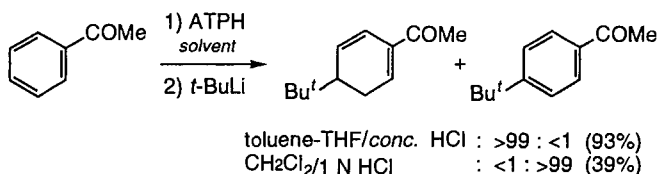
Scheme 2.7

This amphiphilic conjugate alkylation has been successfully applied for the nucleophilic alkylation to electron-deficient arenes based on the unprecedented conjugate addition of organolithiums to aromatic aldehydes and ketones by complexation with ATPH.²⁰ Thus, initial complexation of benzaldehyde or acetophenone with ATPH and subsequent addition of organolithiums affords 1,6-adducts with high selectivity (Scheme 2.8).



Scheme 2.8

The ratio of dearomatization to aromatization products is highly dependent on the choice of solvents and quenching methods as exemplified by the amphiphilic conjugate alkylation to acetophenone (Scheme 2.9).

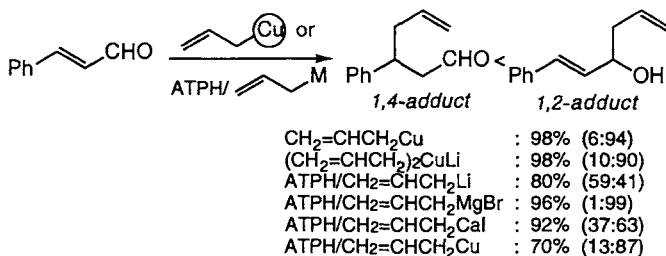


Scheme 2.9

2.3 Amphiphilic conjugate allylations

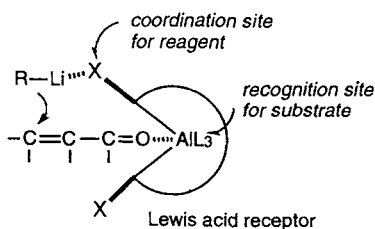
Conjugate allylation to α,β -unsaturated aldehydes is an extremely difficult, hitherto unattainable transformation in organic synthesis, and no effective procedure has yet been developed to a useful level due to the lack of a satisfactory reagent. Even organocopper reagents, which are quite powerful in the conjugate alkylation to α,β -unsaturated carbonyl compounds, gave disappointing results for the conjugate allylation. In fact, attempted reaction of cinnamaldehyde with allylcopper or lithium diallylcuprate gave rise to 1,2-adduct, *trans*-1-phenyl-1,5-hexadien-3-ol predominantly. The new, amphiphilic conjugate alkylation procedure with a Lewis acid receptor, ATPH,¹⁹ was also found to be less effective for the present conjugate allylation, and only the ATPH/allyl-lithium system gave modest 1,4-selectivity (Scheme 2.10).

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

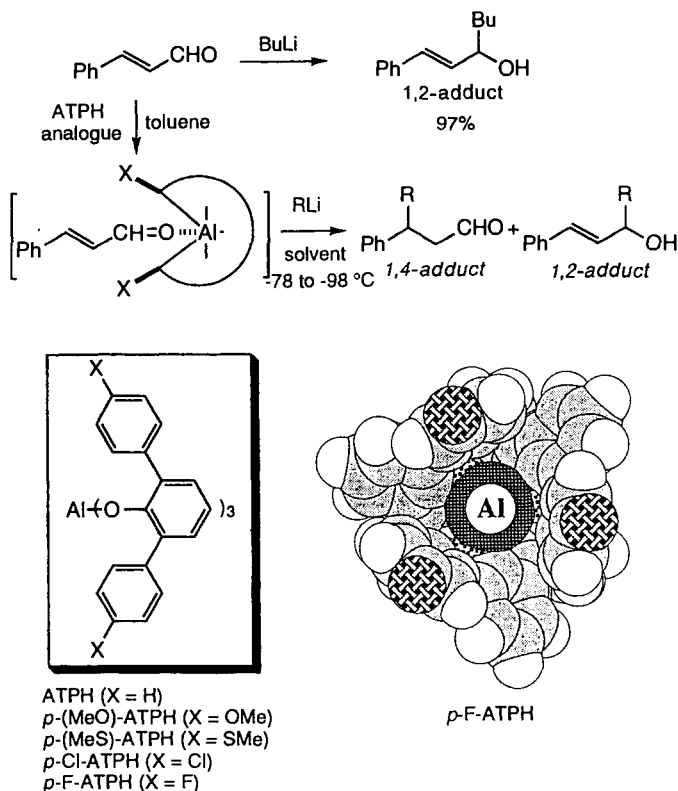


Scheme 2.10

This tendency is contradictory, for example, to our previous observation of the ATPH/Bu-M system for the conjugate alkylation to cinnamaldehyde, where the 1,4-selectivity is enhanced by changing nucleophiles (Bu-M) from BuLi (1.4-/1.2-ratio = 50:50) to BuMgCl (90:10) and BuAl (98:2).¹⁹ Considering the wide availability and versatility of organolithium reagents,²¹⁻²⁴ a new Lewis acid receptor possessing appropriate co-ordination sites for organolithium nucleophiles has been devised (Scheme 2.11). Among various functionalized ATPH derivatives as a Lewis acid receptor, *p*-F-ATPH (Scheme 2.12) was found to be highly effective for this transformation, which clearly demonstrated the synthetic utility of the strong lithium/fluorine participation in selective organic synthesis.²⁵ First, the 1,4-selectivity for the conjugate alkylation to cinnamaldehyde was examined with the modified ATPH/BuLi system in model experiments. Selected results are shown in Table 2.2. *p*-(MeO)-ATPH and *p*-(MeS)-ATPH showed slightly better selectivity than ATPH (Table 2.2, entries 2 and 3). The 1,4-selectivity was further enhanced by designing *p*-Cl-ATPH and *p*-F-ATPH (entries 4 and 5). Significant solvent and temperature effects on the 1,4-selectivity were also observed (entries 6-9), and eventually the optimum reaction condition was achieved using 1,2-dimethoxyethane (DME) solvent for BuLi at lower temperature under the influence of *p*-F-ATPH in toluene, giving the 1,4-adduct with 95% selectivity (entry 9). Here, the chelation of BuLi with DME is quite appropriate to increase the steric size of the nucleophile (BuLi), while still maintaining the co-ordination ability of Li⁺ to fluorine atoms of *p*-F-ATPH. This molecular recognition system is highlighted by the first successful conjugate addition of



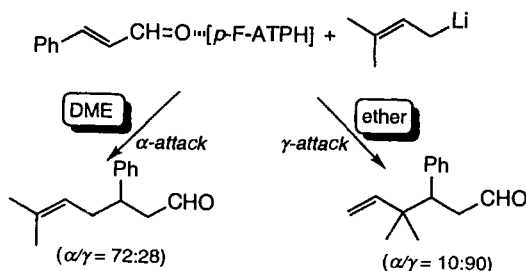
Scheme 2.11



Scheme 2.12

allyl-lithium reagents to α,β -unsaturated aldehydes by complexation with the modified Lewis acid receptor, *p*-F-ATPH (entry 13).

Furthermore, the conjugate addition of prenyl-lithium to cinnamaldehyde proceeded equally well with excellent selectivity under optimized reaction conditions, where the α/γ ratio of the conjugate adducts was profoundly influenced by the solvent effect (Scheme 2.13).



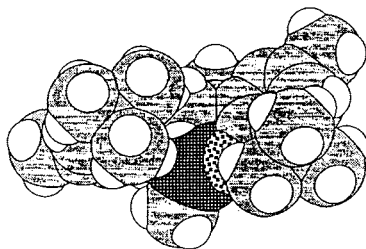
Scheme 2.13

Table 2.2 Conjugate addition of RLi to cinnamaldehyde with modified ATPH

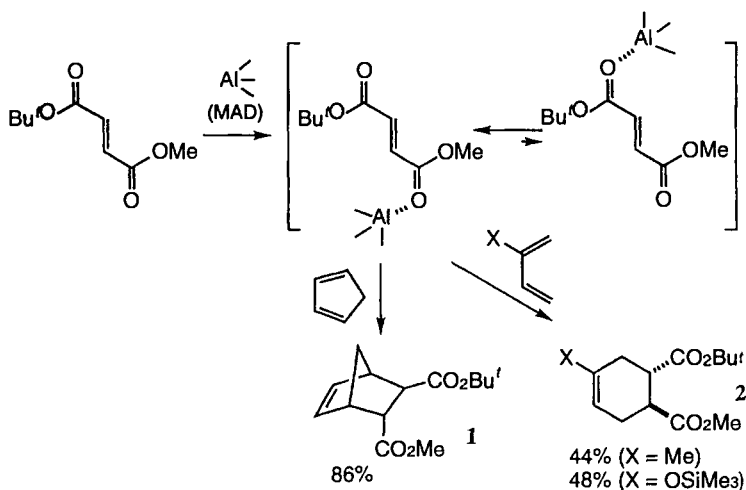
| Entry | ATPH analogue | RLi/solvent | Temp (°C) | Yield (1,4/1,2 ratio) |
|-------|----------------------|----------------|-----------|-----------------------|
| 1 | ATPH | BuLi/hexane | −78 | 92% (50 : 50) |
| 2 | <i>p</i> -(MeO)-ATPH | BuLi/hexane | −78 | 80% (55 : 45) |
| 3 | <i>p</i> -(MeS)-ATPH | BuLi/hexane | −78 | 91% (57 : 43) |
| 4 | <i>p</i> -Cl-ATPH | BuLi/hexane | −78 | 92% (63 : 37) |
| 5 | <i>p</i> -F-ATPH | BuLi/hexane | −78 | 87% (76 : 24) |
| 6 | <i>p</i> -F-ATPH | BuLi/ether | −78 | 90% (79 : 21) |
| 7 | | BuLi/THF | −78 | 82% (86 : 14) |
| 8 | | BuLi/DME | −78 | 75% (90 : 10) |
| 9 | | BuLi/DME | −98 | 83% (95 : 5) |
| 10 | | Allyl-Li-ether | −78 | 94% (77 : 23) |
| 11 | | Allyl-Li/THF | −78 | 89% (50 : 50) |
| 12 | | Allyl-Li/DME | −78 | 75% (90 : 10) |
| 13 | | Allyl-Li-DME | −98 | 83% (95 : 5) |

3. Regio- and stereocontrolled Diels–Alder reaction

The Diels–Alder reaction is undoubtedly the best known and most thoroughly investigated of all cycloaddition reactions because of sustained interest in its mechanism and its exceptionally broad application to regio- and stereo-defined synthesis. The rate of this reaction is usually accelerated by the presence of certain Lewis acids. As revealed by the space-filling model, the exceptionally bulky aluminium reagent, MAD (Scheme 2.14), in addition to its Lewis acidity, provides an exceptionally bulky molecular cleft, which may feature a complementary size, shape, and co-ordination capacity for structurally similar ester substrates. Indeed, MAD allows the discrimination of two structurally different ester carbonyls of unsymmetrical fumarates such as *t*-butyl methyl fumarate, where the sterically less hindered methoxycarbonyl moiety binds selectively to the bulky molecular cleft of MAD as revealed by low-temperature ^{13}C NMR spectroscopy. This finding has been successfully applied to the regio- and stereocontrolled Diels–Alder reaction of unsymmetrical fumarates.²⁶ Thus, the Diels–Alder reaction of the *t*-butyl methyl

**Scheme 2.14.** MAD.

fumarate/MAD complex with cyclopentadiene at -78°C resulted in stereoselective formation of the cycloadduct **1** (Scheme 2.15) almost exclusively. In addition, treatment of the 1:1 co-ordination complex with 2-substituted 1,3-butadiene ($\text{X} = \text{Me}$, OSiMe_3) gave the cycloadduct **2** ($\text{X} = \text{Me}$, OSiMe_3) with high regioselectivity. In marked contrast, the cycloadditions with Et_2AlCl as an ordinary Lewis acid were found to have a total lack of selectivity.

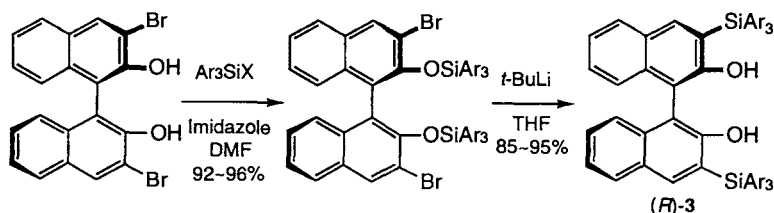
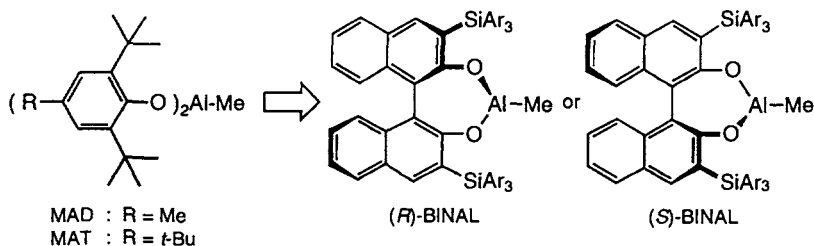


Scheme 2.15

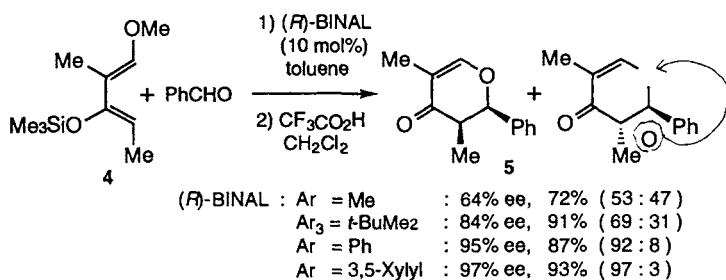
4. Asymmetric hetero-Diels–Alder reaction

Based on the concept of the diastereoselective activation of carbonyl groups with MAD or MAT as described in the section 2.1, the first reliable, bulky chiral organoaluminium reagent, (*R*)-BINAL or (*S*)-BINAL has been devised for enantioselective activation of carbonyl groups. The sterically hindered, enantiomerically pure (*R*)-(+)-3,3'-bis(triarylsilyl)binaphthol((*R*)-**3**) required for the preparation of (*R*)-BINAL can be synthesized in two steps from (*R*)-(+)-3,3'-dibromobinaphthol by bis-triarylsilylation and subsequent intramolecular 1,3-rearrangement of the triarylsilyl groups as shown in Scheme 2.16.²⁷ Reaction of (*R*)-**3** in toluene with trimethylaluminium produced the chiral organoaluminium reagent (*R*)-BINAL quantitatively. Its molecular weight, found cryoscopically in benzene, corresponds closely to the value calculated for the monomeric species. The modified chiral organoaluminium reagents, (*R*)-BINAL and (*S*)-BINAL were shown to be highly effective as chiral Lewis acid catalysts in the asymmetric hetero Diels–Alder reaction.²⁸ Reaction of various aldehydes with activated diene **4** (Scheme 2.17) under the influence of a catalytic amount of BINAL (5–10 mol%) at -20°C , after ex-

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors



Scheme 2.16



Scheme 2.17

posure of the resulting hetero Diels–Alder adducts to trifluoroacetic acid, gave predominantly *cis*-dihydropyrene **5** in high yield with excellent enantioselectivity. The enantioface differentiation of prochiral aldehydes is controllable by judicious choice of the size of trialkylsilyl moiety in BINAL, thereby allowing the rational design of the catalyst for asymmetric induction. In fact, switching the triarylsilyl substituent (Ar = Ph or 3,5-xylyl) to the *tert*-butyldimethylsilyl or trimethylsilyl group led to a substantial loss of enantio as well as *cis* selectivity in the hetero Diels–Alder reaction of benzaldehyde and activated diene **4**. In marked contrast, the chiral organoaluminium reagent derived from trimethylaluminium and (*R*)-(+)-3,3'-dialkylbinaphthol (alkyl = H, Me, or Ph) could be utilized, but only as a stoichiometric reagent and results were disappointing both in terms of reactivity and enantioselectivity for this hetero Diels–Alder reaction.

Protocol 2.**Synthesis of (5*R*, 6*R*)-3,5-dimethyl-6-phenyldihydropyran-4-one (5).
Asymmetric hetero-Diels–Alder reaction catalysed by (*R*)-BINAL**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles. Because of its high oxygen and moisture sensitivity, (*R*)-BINAL reagent should be prepared *in situ* and used immediately.

Equipment

- Magnetic stirrer
- Three-necked, round-bottomed flask (200 mL).
A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Teflon-coated magnetic stirring bar
- Medium-gauge needle
- All-glass syringe with a needle-lock Luer (volume appropriate for quantity of solution to be transferred)
- Vacuum/inert gas source (argon or nitrogen source may be an argon or a nitrogen balloon)

Materials

- (*R*)-(+)-3,3'-Bis(triphenylsilyl)binaphthol (FW 800), 0.88 g, 1.1 mmol irritant
- Trimethylaluminum in hexane (FW 72.1), 1 M solution in hexane, 1 mL, 1 mmol pyrophoric, moisture sensitive
- Dry toluene, 50 mL flammable, toxic
- Benzaldehyde (FW 106.1), 1.06 g, 10 mmol highly toxic, cancer suspecting agent
- (1*E*,3*Z*)-2,4-Dimethyl-1-methoxy-3-(trimethylsiloxy)-1,3-butadiene (FW 200.4), 2.20 g, 11 mmol moisture sensitive
- 10% HCl, 50 mL toxic
- 1 M NaHCO₃, 50 mL
- Technical ether for extraction, 50 mL flammable, toxic
- Technical dichloromethane, 100 mL toxic, irritant
- Trifluoroacetic acid (FW 114.0), 1.37 g, 12 mmol corrosive, toxic
- Silica gel for flash chromatography, 200 g, Merck Kieselgel 60 (Art. 9385) irritant dust
- Ether for flash chromatography flammable, toxic
- *n*-Hexane for flash chromatography flammable, irritant

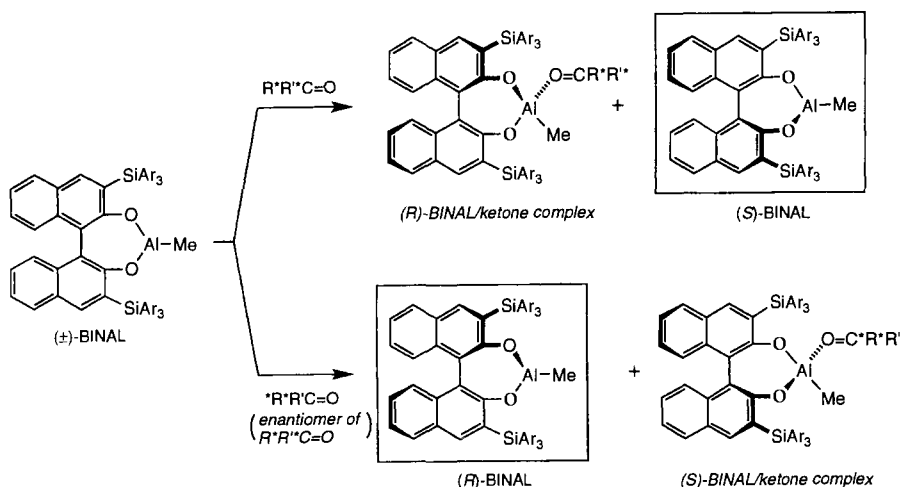
1. Clean all glassware, syringes, needles, and stirring bar and dry at least 2 h in an electric oven at 100°C before use.
2. Assemble the flask, stirring bar, and stop cocks under argon while the apparatus is still hot.
3. Support the assembled flask using a clamp and a stand with a heavy base.
4. Dry the apparatus with an electric heat gun under vacuum (1–2 mm Hg) for 5 min, then back-fill the flask with argon. Repeat to a total of three times.
5. Place (*R*)-(+)-3,3'-bis(triphenylsilyl)binaphthol in the flask and flush with argon.
6. Add dry toluene (50 mL).
7. Stir the mixture, degassed under vacuum, and replaced by argon.

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

8. Support the bottle containing trimethylaluminium in hexane using a clamp and a stand with a heavy base.
9. Fill a syringe with trimethylaluminium in hexane from the bottle containing trimethylaluminium using argon pressure. Apply the argon pressure to fill the syringe slowly with the required volume. Transfer the reagent in the syringe to the reaction flask at room temperature.
10. Stir the resulting solution at room temperature for 1 h to give (*R*)-BINAL (Ar = Ph) almost quantitatively. During this operation, nearly 2 equivalents of methane gas are evolved per 1 equiv of trimethylaluminium.
11. Cool the reaction vessel to a temperature of -20°C in a dry ice/*o*-xylene bath. *o*-Xylene is recommended as refrigerant in place of carbon tetrachloride (toxic and cancer suspecting agent).
12. Add benzaldehyde and (1*E*,3*Z*)-2,4-dimethyl-1-methoxy-3-(trimethylsiloxy)-1,3-butadiene sequentially at -20°C . Stir the mixture at -20°C for 2 h in order to complete the cycloaddition.
13. Place 10% HCl solution and the stirring bar in the Erlenmeyer flask. While stirring the diluted HCl solution vigorously, add the reaction mixture slowly at 0°C to avoid excessive foaming on hydrolysis.
14. Remove the ice bath and stir the entire mixture vigorously at 25°C for 20 min.
15. Transfer the mixture to a separating funnel and separate the two layers. Extract the water layer with ether twice ($2 \times 25\text{ mL}$).
16. Combine the ethereal extracts to a 300 mL flask. Dry the layer over anhydrous magnesium sulfate, and filter through a filter paper. Concentrate the filtrate under reduced pressure by means of a rotary evaporator.
17. Dilute the oily residue with dichloromethane.
18. Add trifluoroacetic acid at 0°C , and stir the mixture at 0°C for 1 h.
19. Transfer the mixture to a separating funnel containing diluted Na/CO₃ solution, and separate the two layers. Extract the water layer with dichloromethane twice ($2 \times 25\text{ mL}$).
20. Combine the extracts in a flask. Dry the layer over anhydrous magnesium sulfate, and filter through a filter paper. Concentrate the filtrate under reduced pressure by means of a rotary evaporator.
21. Purify the oily residue by column chromatography on silica gel (ether/hexane = 1:3 as eluents) to give 1.56 g (77%, 95% ee) of (5*R*,6*R*)-3,5-dimethyl-6-phenyldihydropyran-4-one which is characterized by ¹H NMR analysis: $[\alpha]_{\text{D}} +7.1^{\circ}$ (*c* 1.0, chloroform).

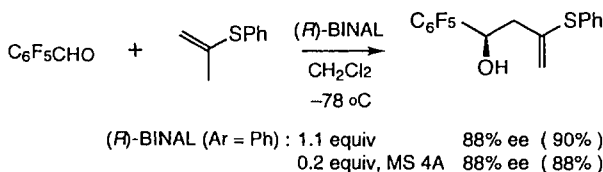
An interesting method for the preparation of chiral aluminium reagents has appeared recently. The chiral organoaluminium reagent, (*R*)-BINAL or (*S*)-BINAL can be generated *in situ* from the corresponding racemate (\pm)-BINAL by diastereoselective complexation with certain chiral ketones

(Scheme 2.18).²⁹ Among several terpene-derived chiral ketones, 3-bromocamphor was found to be the most satisfactory. The hetero Diels–Alder reaction of benzaldehyde and 2,4-dimethyl-1-methoxy-3-trimethylsiloxy-1,3-butadiene (**4**) with 0.1 equiv. of (\pm)-BINAL (Ar = Ph) and *d*-bromocamphor at -78°C gave rise to *cis*-adduct **5** as the major product with 82% ee. Although the level of asymmetric induction attained does not yet match that acquired with the enantiomerically pure BINAL (Ar = Ph, 95% ee), one recrystallization of the *cis*-adduct **5** of 82% ee from hexane gave essentially enantiomerically pure **5**, thereby enhancing the practicality of this method. This study demonstrates the potential for broad application of the *in situ* generated chiral catalyst via diastereoselective complexation in asymmetric synthesis.



Scheme 2.18

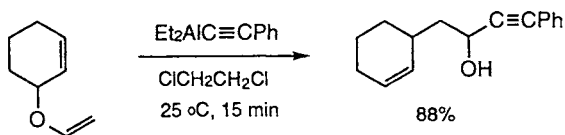
Since the enantioselective activation of carbonyl with the chiral aluminium, (*R*)-BINAL or (*S*)-BINAL, had been demonstrated, the asymmetric ene reaction of electron-deficient aldehydes with various alkenes, by using the latter reagent, could also be considered a feasible transformation.³⁰ Indeed in the presence of powdered 4Å molecular sieves, the chiral aluminium reagent, (*R*)-BINAL or (*S*)-BINAL can be used as a catalyst without any loss of enantioselectivity (Scheme 2.19).



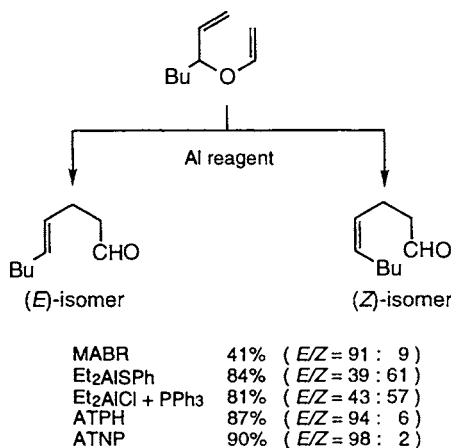
Scheme 2.19

5. Claisen rearrangement

Aliphatic Claisen rearrangements normally require high temperatures. However, in the presence of Lewis acidic organoaluminium reagents, the rearrangements have been accomplished under very mild conditions. Treatment of simple allyl vinyl ether substrates with trialkylaluminums resulted in the [3,3] sigmatropic rearrangement and subsequent alkylation on the aldehyde carbonyl group³¹ (Scheme 2.20). Notably, ordinary strong Lewis acids such as TiCl_4 , SnCl_4 , and $\text{BF}_3 \cdot \text{OEt}_2$ could not be used for this rearrangement. The rearrangement–reduction product was obtained exclusively with $i\text{-Bu}_3\text{Al}$ or DIBAH. The aluminium thiolate, Et_2AlSPh or a combination of Et_2AlCl and PPh_3 was effective for the rearrangement providing the normal Claisen products, γ,δ -unsaturated aldehydes, however, without any stereoselectivity. Accordingly, a new molecular recognition approach for the stereocontrolled Claisen rearrangement of allyl vinyl ethers has been developed based on the stereoselective activation of ether moiety using aluminium-type Lewis acid receptors. Thus, treatment of 1-butyl-2-propenyl vinyl ether with ATPH in CH_2Cl_2 afforded *E*-Claisen products predominantly (*E/Z* ratio = 94:6) (Scheme 2.21). Use of sterically more hindered aluminium tris(2- α -naphthyl-6-phenylphenoxide) (ATNP) exhibited better selectivity (*E/Z* ratio = 98:2).³²



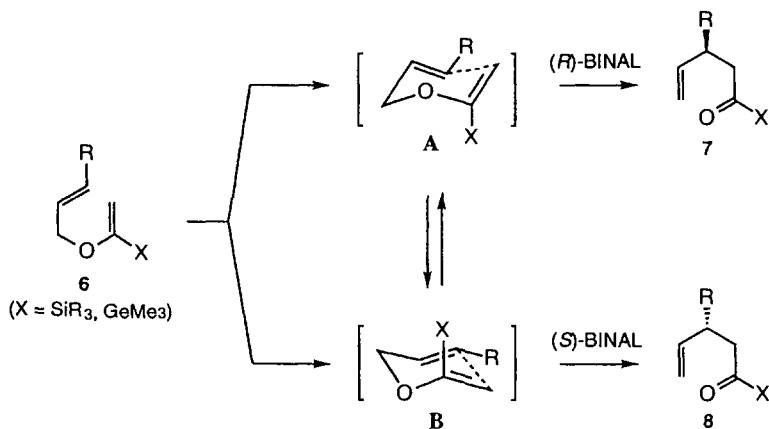
Scheme 2.20



Scheme 2.21

In marked contrast, use of exceptionally bulky, Lewis acidic receptor, methylaluminium bis(2,6-di-*tert*-butyl-4-bromophenoxide) (MABR) resulted in predominant formation of *Z*-Claisen products (*E/Z* ratio = 9:91), which was very difficult to attain by the conventional methodologies including thermal Claisen rearrangement and its variants (Carroll, the ortho ester, Eschenmoser, and Ireland rearrangements).³³

In a similar manner, the concept of the enantioselective activation of carbonyl groups with the bulky, chiral aluminums, (*R*)-BINAL or (*S*)-BINAL, has also been extended to the enantioselective activation of an ether oxygen, which gave rise to the first successful example of the asymmetric Claisen rearrangement of allylic vinyl ethers **6** catalyzed by (*R*)-BINAL or (*S*)-BINAL^{34,35} (Scheme 2.22). This method provides an easy asymmetric synthesis of various acylsilanes **7** or **8** ($X = \text{SiR}_3$) and acylgermane **7** ($X = \text{GeMe}_3$) with high enantiomeric purity (Table 2.3). Among the various trialkylsilyl substituents



Scheme 2.22

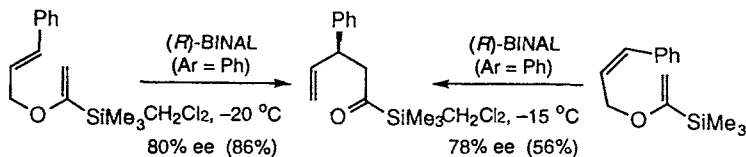
Table 2.3 Asymmetric Claisen rearrangement of allylic vinyl ether **6** (Scheme 2.22)

| Allyl vinyl ether | (<i>R</i>)-BINAL | Product | Yield | Optical yield |
|--|--|----------|-------|---------------|
| 6 (R=Ph, X=SiMe ₃) | Ar ₃ =Bu ^t Me ₂ | 7 | 22% | 14% ee |
| 6 (R=Ph, X=SiMe ₃) | Ar=Ph | 7 | 86% | 80% ee |
| 6 (R=Ph, X=SiMe ₃) | Ar ₃ =Bu ^t Ph ₂ | 7 | 99% | 88% ee |
| 6 (R=Ph, X=SiMe ₃) | Ar=Ph ^a | 8 | 85% | 80% ee |
| 6 (R=Ph, X=SiMe ₂ Ph) | Ar=Ph | 7 | 65% | 85% ee |
| 6 (R=Ph, X=SiMe ₂ Ph) | Ar ₃ =Bu ^t Ph ₂ | 7 | 76% | 90% ee |
| 6 (R=cyclohexyl X=SiMe ₃) | Ar=Ph | 7 | 79% | 61% ee |
| 6 (R=cyclohexyl X=SiMe ₃) | Ar ₃ =Bu ^t Ph ₂ | 7 | 84% | 71% ee |
| 6 (R=Ph, X=GeMe ₃) | Ar=Ph | 7 | 73% | 91% ee |
| 6 (R=Ph, X=GeMe ₃) | Ar ₃ =Bu ^t Ph ₂ | 7 | 68% | 93% ee |

^aUse of (*S*)-BINAL as catalyst.

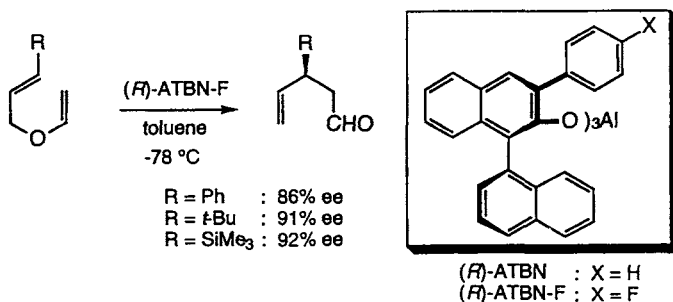
of BINAL, use of a bulkier *t*-butyldiphenylsilyl group gives rise to the highest enantioselectivity. Conformational analysis of two possible chairlike transition-state structures of an allyl vinyl ether substrate **6** reveals that the chiral organoaluminium reagent BINAL can discriminate between these two conformers **A** and **B**, which differ from each other only in the orientation of α -methylene groups of ethers.

Notably, the asymmetric Claisen rearrangement of *cis*-allylic α -(trimethylsilyl)vinyl ethers with (*R*)-BINAL produced optically active acylsilanes with the same absolute configuration as those produced from *trans*-allylic α -(trimethylsilyl)vinyl ethers³⁶ (Scheme 2.23).



Scheme 2.23

The bulky, chiral organoaluminium reagents of type (*R*)-BINAL or (*S*)-BINAL is only applicable for allyl vinyl ether substrates possessing bulky α -silyl and α -germyl substituents. Later, the asymmetric Claisen rearrangement of simple allyl vinyl ether substrates has been developed with the designing of chiral ATPH analogues, aluminium tris((*R*)-1- α -naphthyl-3-phenylnaphthoxide) ((*R*)-ATBN) or aluminium tris((*R*)-3-(*p*-fluorophenyl)-1- α -naphthyl-naphthoxide) ((*R*)-ATBN-F) with high enantioselectivity³² (Scheme 2.24).

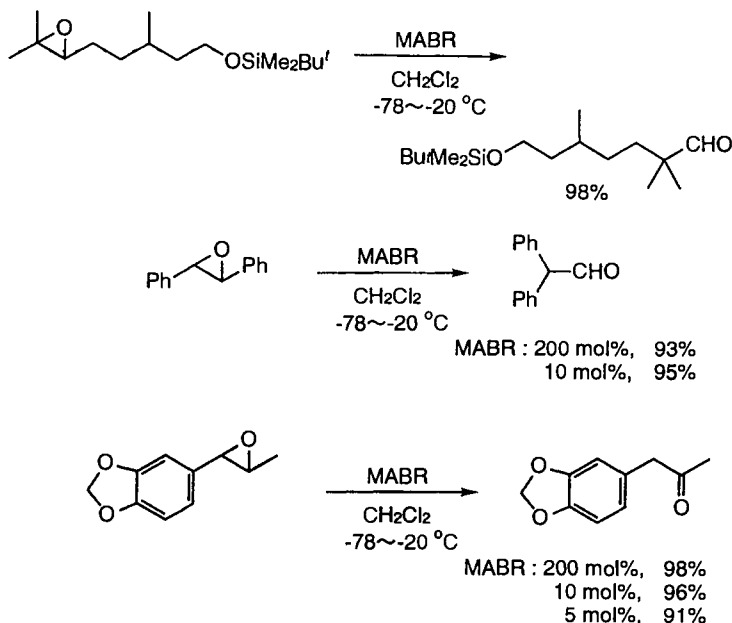


Scheme 2.24

6. Epoxide rearrangement

The exceptional bulkiness of the modified organoaluminium reagent, MABR, can also be utilized for the rearrangement of epoxy substrates under very mild conditions with high efficiency and selectivity. So far, $\text{BF}_3 \cdot \text{OEt}_2$ catalyst is regarded as a reliable Lewis acid catalyst for the epoxide rearrangement.

However, attempted rearrangement of *tert*-butyldimethylsilyl ether of epoxy citronellol with $\text{BF}_3 \cdot \text{OEt}_2$ resulted in the formation of a number of products. In contrast, treatment of this substrate with MABR under mild conditions (-78 to -20°C) gave the desired aldehyde almost quantitatively (Scheme 2.25). In addition, certain epoxy substrates can be rearranged with catalytic use of MABR.³⁷



Scheme 2.25

Protocol 3.

Synthesis of diphenylacetaldehyde. MABR-catalysed rearrangement of *trans*-stilbene oxide

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles. Because of its high oxygen and moisture sensitivity, MABR reagent should be prepared *in situ* and used immediately.

Equipment

- Magnetic stirrer
- All-glass syringe with a needle-lock Luer (volume appropriate for quantity of solution to be transferred)
- Teflon-coated magnetic stirrer bar
- Vacuum/inert gas source (argon source may be an argon balloon)
- Medium-gauge needle
- Three-necked, round-bottomed flask (1 L) fitted with a condenser and a pressure-equalizing dropping funnel. A three-way stopcock is fitted to the top of the condenser and connected to a vacuum/argon source

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

Materials

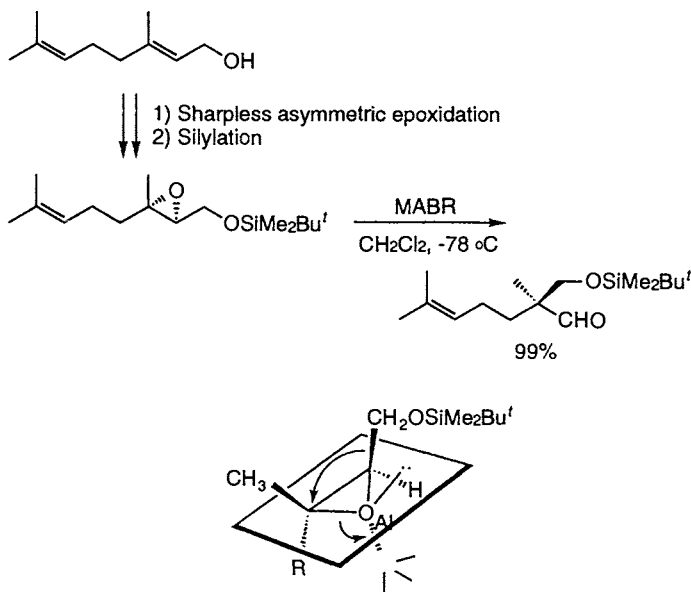
- 4-Bromo-2,6-di-*tert*-butylphenol (FW 285.2), 3.42 g, 12 mmol **irritant**
- Trimethylaluminium in hexane (FW 72.1), 2 M solution in hexane, 3 mL, 6 mmol **pyrophoric, moisture sensitive**
- Dry dichloromethane, 300 mL **toxic, irritant**
- *trans*-Stilbene oxide (FW), 11.78 g, 60 mmol
- Sodium fluoride (FW 42.0), 1.01 g, 24 mmol **moisture sensitive, toxic**
- Water (FW 18), 324 μ L, 18 mmol
- Technical dichloromethane **toxic, irritant**
- Silica gel for flash chromatography, 500 g, Merck Kieselgel 60 (Art. 9385) **irritant dust**
- Ether for flash chromatography **flammable, toxic**
- Dichloromethane for flash chromatography **toxic, irritant**
- *n*-Hexane for flash chromatography **flammable, irritant**

1. Clean all glassware, syringes, needles, and stirring bar and dry for at least 2 h in an electric oven at 100°C before use.
2. Assemble the flask, stirring bar, and stop cocks under argon while the apparatus is still hot.
3. Support the assembled flask using a clamp and a stand with a heavy base.
4. Dry the apparatus with an electric heat gun under vacuum (1–2 mm Hg) for 5 min, then back-fill the flask with argon. Repeat to a total of three times.
5. Place 4-bromo-2,6-di-*tert*-butylphenol and flush with argon.
6. Add freshly distilled dichloromethane (300 mL).
7. Stir the mixture, degas under vacuum, and replace by argon.
8. Support the bottle containing trimethylaluminium in hexane using a clamp and a stand with a heavy base.
9. Fill a syringe with trimethylaluminium in hexane from the bottle containing trimethylaluminium using argon pressure. Apply the argon pressure to fill the syringe slowly with the required volume. Transfer the reagent in the syringe to the reaction flask at room temperature.
10. Stir the resulting solution at this temperature for 1 h to give methylaluminium bis(4-bromo-2,6-di-*tert*-butylphenoxide) (MABR) almost quantitatively. During this operation, nearly 2 equivalents of methane gas are evolved per 1 equiv of trimethylaluminium.
11. Cool the reaction vessel to –20°C in a dry ice/*o*-xylene bath. *o*-Xylene is recommended as refrigerant in place of carbon tetrachloride (toxic and cancer suspecting agent).
12. Transfer *trans*-stilbene oxide, which is dissolved in 25 mL of dry dichloromethane, to the dropping funnel, and added to the reaction flask over 15–20 min at –20°C. Stir the mixture at –20°C for 20 min in order to complete the rearrangement.
13. Add sodium fluoride, and injected water dropwise at –20°C. To avoid excessive foaming on hydrolysis water should be added carefully by syringe.

Protocol 3. Continued

14. Stir the entire mixture vigorously at -20°C for 5 min at 0°C for 30 min.
15. Filter the contents of the flask with the aid of three 50 mL portions of dichloromethane. The sodium fluoride–water workup offers an excellent method for large-scale preparations, and is generally applicable for product isolation in the reaction of organoaluminium compounds.
16. Evaporate the combined filtrates under reduced pressure with a rotary evaporator. Purify the oily residue by column chromatography (column diameter: 9.5 cm) on silica gel (ether/dichloromethane/hexane = 1:2:20 to 1:1:10 as eluants) to give 11.02–11.17 g (94–95%) of diphenylacetaldehyde as a colourless oil, which is characterized by ^1H NMR analysis.


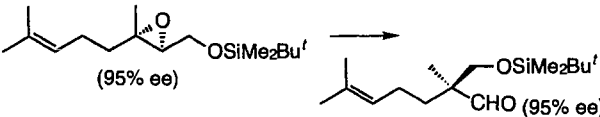
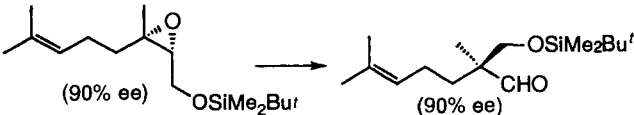

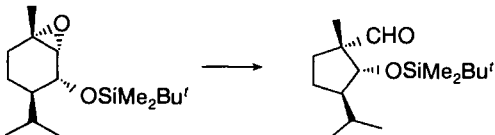
Although the acid-catalysed rearrangement of epoxides to carbonyl compounds is a well-known transformation and a number of reagents have been elaborated for this purpose, only a few reagents have been used successfully for the rearrangement of functionalized epoxides with respect to the efficiency and selectivity of the reaction. With the stoichiometric use of MABR, however, a new, stereocontrolled rearrangement of epoxy silyl ethers leading to β -siloxy aldehydes has been developed under mild conditions. Interestingly, used in combination with the Sharpless asymmetric epoxidation of allylic alcohols, this rearrangement represents a new approach to the synthesis of various optically active β -hydroxy aldehydes,^{38,39} (Scheme 2.26) which are

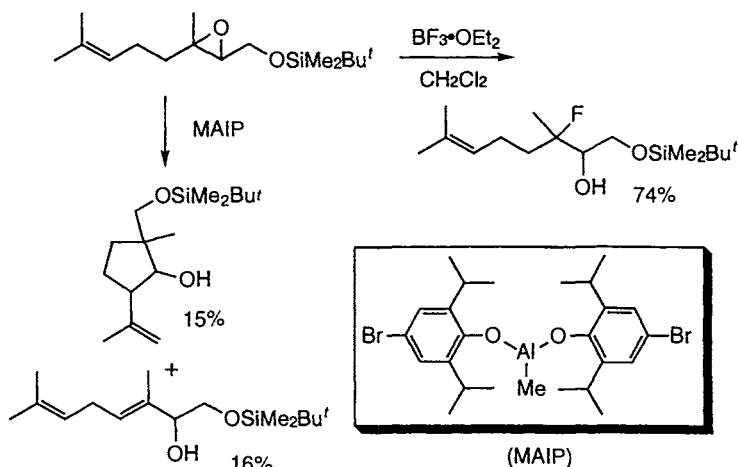
**Scheme 2.26**

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

quite useful intermediates in natural product synthesis. Based on the optical rotation sign and value of the β -siloxy aldehydes, this organoaluminium-promoted rearrangement proceeds with rigorous transfer of the epoxide chirality, and the observed stereoselectivity can be interpreted to arise from the *anti* migration of the siloxymethyl group to the epoxide moiety. Here

Table 2.4 MABR-catalysed rearrangement of optically active epoxy silyl ethers

| Entry | Mol % of MABR | Conditions (°C, h) | Chemical yield |
|--|---------------|----------------------------|----------------|
|  <p>(98% ee) (98% ee)</p> | | | |
| 1 | 200 | -78, 0.5 | 95% |
| 2 | 20 | -20, 0.3 | 82% |
| 3 | 10 | -20, 0.3 | 74% |
|  <p>(95% ee) (95% ee)</p> | | | |
| 4 | 200 | -78, 1 | 99% |
| 5 | 20 | -78, 0.2; 0, 0.5 | 82% |
| 6 | 10 | -78, 0.2; 0, 1 | 74% |
|  <p>(90% ee) (90% ee)</p> | | | |
| 7 | 200 | -78, 1; -40, 0.5 | 98% |
| 8 | 20 | -78, 0.2; 0, 1 | 79% |
| 9 | 10 | -78, 0.2; 0, 3 | 68% |
|  <p>(>98% ee) (>98% ee)</p> | | | |
| 10 | 200 | -78, 0.5 | 87% |
| 11 | 20 | -40, 0.5 | 75% |
| 12 | 10 | -78, 0.2; -20, 0.5, 0, 0.5 | 71% |
|  | | | |
| 13 | 200 | -78, 0.5; -20, 2 | 88% |
| 14 | 20 | -78, 0.2; 0, 5 | 77% |



Scheme 2.26

again, the exceptional bulkiness of 2,6-di-*tert*-butyl-4-bromophenoxy ligands in MABR is essential for the smooth rearrangement of epoxy silyl ethers, and the less bulky methylaluminum bis(4-bromo-2,6-di-isopropylphenoxide) (MAIP) was found to be totally ineffective for the rearrangement of *tert*-butyldimethylsilyl ether of epoxygeraniol (Table 2.4). Again, $\text{BF}_3 \cdot \text{OEt}_2$ as an ordinary Lewis acid gave fluorohydrines as sole isolable products (Scheme 2.27).

References

1. Mole, T.; Jeffery, E. A. *Organoaluminum compounds*. Elsevier, Amsterdam **1972**.
2. Reinheckel, H.; Haage, K.; Jahnke, D. *Organometal. Chem. Rev. A* **1969**, *4*, 47.
3. Lehmkuhl, H.; Ziegler, K.; Gellert, H.; Houben-Weyl, G. *Methoden der Organischen Chemie* (4th edn) Vol. XIII, Part 4. Thieme, Stuttgart, **1970**.
4. Bruno, G. *The use of aliminum alkyls in organic synthesis*. Ethyl Corporation, Baton Rouge, FL, USA, **1970, 1973, 1980**.
5. Negishi, E. *J. Organometal. Chem. Libr.* **1982**, *1*, 93.
6. Yamamoto, H.; Nozaki, H. *Angew. Chem. Int. Ed. Engl.* **1978**, *17*, 169.
7. Negishi, E. *Organometallics in organic synthesis*. Vol. 1, Wiley, New York, **1980**, p. 286.
8. Eisch, J. J. In: *Comprehensive organometallic chemistry* (ed. G. Wilkinson, F. G. A. Stone, E. W. Abel). Pergamon Press, Oxford, **1982**, Vol. 1, p 555.
9. Zietz, Jr, J. R.; Robinson, G. C.; Leindsay, K. L. In: *Comprehensive organometallic chemistry* (ed. G. Wilkinson, F. G. A. Stone, E. W. Abel). Pergamon Press, Oxford, **1982**, Vol. 7, p 365.
10. Zweifel, G.; Miller, J. A. *Org. React.* **1984**, *32*, 375.
11. Maruoka, K.; Yamamoto, H. *Angew. Chem. Int. Ed. Engl.* **1985**, *24*, 668.

2: Synthetic use of bulky aluminium reagents as Lewis acid receptors

12. Maruoka, K.; Yamamoto, H. *Tetrahedron* **1988**, *44*, 5001.
13. Maruoka, K. In: *Synthesis of Organometallic Compounds* (ed. S. Komiya). Wiley, Chichester, **1997**.
14. Ziegler, K. In: *Organometallic Chemistry* (ed. H. Zeiss). Reinhold, New York, **1960**, p 194.
15. Wilke, G. In: *Coordination Polymerization* (ed. J. C. W. Chien). Academic Press, New York, **1975**.
16. Strarowieyski, K. B.; Pasynkiewics, S.; Skowronska-Ptasinska, M. *J. Organometal. Chem.* **1975**, *90*, C43.
17. Maruoka, K.; Itoh, T.; Yamamoto, H. *J. Am. Chem. Soc.* **1985**, *107*, 4573.
18. Maruoka, K.; Sakurai, M.; Yamamoto, H. *Tetrahedron Lett.* **1985**, *26*, 3853.
19. Maruoka, K.; Imoto, H.; Saito, S.; Yamamoto, H. *J. Am. Chem. Soc.* **1994**, *116*, 4131.
20. Maruoka, K.; Ito, M.; Yamamoto, H. *J. Am. Chem. Soc.*, **1995**, *117*, 9091.
21. Wakefield, B. J. *The chemistry of organolithium compounds*. Pergamon, Oxford, **1974**.
22. Beswick, M. A.; Wright, D. S. In: *Comprehensive organometallic chemistry II* (ed. E. W. Abel, F. G. A. Stone, G. Wilkinson). Pergamon, Oxford, **1995**, Vol. 1, 1.
23. Gray, M.; Tinkl, M.; Snieckus, V. In: *Comprehensive organometallic chemistry II* (ed. E. W. Abel, F. G. A. Stone, G. Wilkinson). Pergamon, Oxford, **1995**, Vol. 11, 1.
24. Wakefield, B. J. *Organolithium Method*. Academic Press, London, **1988**.
25. Ooi, T.; Kondo, Y.; Maruoka, K. *Angew. Chem. Int. Ed. Engl.*, **1997**, *36*, 1183.
26. Maruoka, K.; Saito, S.; Yamamoto, H. *J. Am. Chem. Soc.* **1992**, *114*, 1089.
27. Maruoka, K.; Ito, T.; Araki, Y.; Shirasaka, T.; Yamamoto, H. *Bull. Chem. Soc. Jpn* **1988**, *61*, 2975.
28. Maruoka, K.; Ito, T.; Shirasaka, T.; Yamamoto, H. *J. Am. Chem. Soc.* **1988**, *110*, 310.
29. Maruoka, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1989**, *111*, 789.
30. Maruoka, K.; Hoshino, Y.; Shirasaka, T.; Yamamoto, H. *Tetrahedron Lett.*, **1988**, *29*, 3967.
31. Takai, K.; Mori, I.; Oshima, K.; Nozaki, H. *Bull. Chem. Soc., Jpn* **1986**, *57*, 446.
32. Maruoka, K.; Saito, S.; Yamamoto, H. *J. Am. Chem. Soc.* **1995**, *117*, 6153.
33. Nonoshita, K.; Banno, H.; Maruoka, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1990**, *112*, 316.
34. Maruoka, K.; Banno, H.; Yamamoto, H. *J. Am. Chem. Soc.* **1990**, *112*, 7791.
35. Maruoka, K.; Banno, H.; Yamamoto, H. *Tetrahedron-Asymmetry* **1991**, *2*, 647.
36. Maruoka, K.; Yamamoto, H. *Synlett* **1991**, 793.
37. Maruoka, K.; Nagahara, S.; Ooi, T.; Yamamoto, H. *Tetrahedron Lett.* **1989**, *30*, 5607.
38. Maruoka, K.; Ooi, T.; Yamamoto, H. *J. Am. Chem. Soc.* **1989**, *111*, 6431.
39. Maruoka, K.; Ooi, T.; Nagahara, S.; Yamamoto, H. *Tetrahedron* **1991**, *47*, 6983.

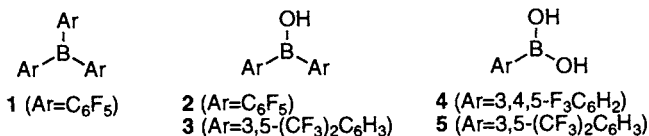
Boron reagents

KAZUAKI ISHIHARA

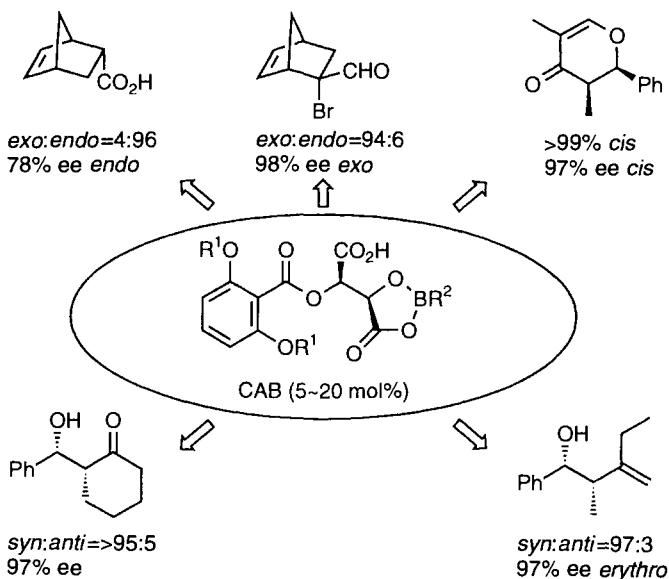
1. Introduction

Arylboron reagents with electron-withdrawing aromatic groups and chiral boron reagents as Lewis acid catalysts are described in this chapter.

The classical boron Lewis acids, BX_3 , RBX_2 and R_2BX ($\text{X} = \text{F}, \text{Cl}, \text{Br}, \text{OTf}$) have become popular tools in organic synthesis. In general, these are used stoichiometrically in organic transformations under anhydrous conditions, since the presence of even a small amount of water causes rapid decomposition or deactivation of the promoters. To obviate some of these inherent problems, we have demonstrated the potential of tris(pentafluorophenyl)boron (**1**), bis(pentafluorophenyl)boronic acid (**2**), bis(3,5-bis(trifluoromethyl)phenyl)boronic acid (**3**), 3,4,5-trifluorophenylboronic acid (**4**), and 3,5-bis(trifluoromethyl)phenylboronic acid (**5**) as a new class of boron catalysts.



Chiral boron catalysts have been widely used as Lewis acids in the asymmetric Diels–Alder, Mukaiyama aldol, Sakurai–Hosomi allylation, and aldol-type reactions of imines. As one successful example, we have achieved highly enantioselective carbo-Diels–Alder, hetero-Diels–Alder, aldol, and allylation reactions using a common chiral (acyloxy)borane (CAB) catalyst as depicted in Scheme 3.1. These carbon–carbon bond formations are very important and useful in asymmetric synthesis.



Scheme 3.1

2. Arylboron reagents with electron-withdrawing aromatic group

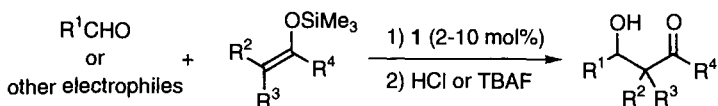
2.1 Tris(pentafluorophenyl)boron as an efficient, air stable, and water-tolerant Lewis acid catalyst

1 is an air stable and water-tolerant Lewis acid catalyst, which is readily prepared as a white solid from boron trichloride by reaction with pentafluorophenyl-lithium.^{1,2} This reagent does not react with pure oxygen.^{1,2} It is very thermally stable, even at 270°C, and soluble in many organic solvents.^{1,2} Although **1** is available when exposed to air (not anhydrous grade), it acts better as a catalyst under anhydrous conditions.

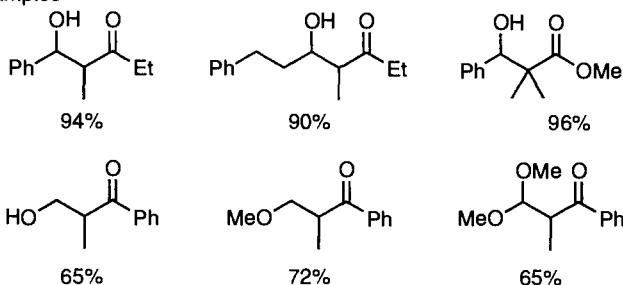
2.1.1 Mukaiyama aldol reactions of silyl enol ethers with aldehydes or other electrophiles^{3,5}

The aldol-type reactions of various silyl enol ethers with aldehydes or other electrophiles proceed smoothly in the presence of 2–10 mol% of **1**. Although the reaction is remarkably promoted by using an anhydrous solution of **1**, the reaction of silyl enol ethers with a commercial aqueous solution of formaldehyde takes place without incident. Silyl enol ethers react with chloromethyl methyl ether or trimethylorthoformate; hydroxymethyl, methoxymethyl, or dimethoxymethyl C1 groups can be introduced at the α -position of the carbonyl group (Scheme 3.2).

3: Boron reagents



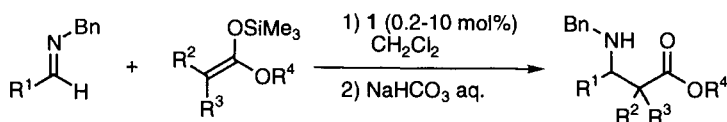
Examples



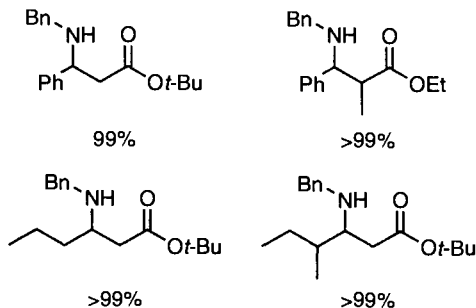
Scheme 3.2

2.1.2 Aldol-type reaction of ketene silyl acetals with *N*-benzylimines^{4,5}

1 (anhydrous grade) is a highly active catalyst for the aldol-type reaction between ketene silyl acetals and *N*-benzylimines because of its stability and comparatively low value of bond energy and affinity toward nitrogen-containing compounds (Scheme 3.3). *N*-Benzylimines are useful substrates because β -benzylamino acid esters produced are readily debenzylated by hydrogenolysis over palladium on carbon. Catalysis is carried out using 0.2–10 mol% catalyst loading in toluene. The condensation proceeds smoothly even with aliphatic enolizable imines derived from primary and secondary aliphatic



Examples



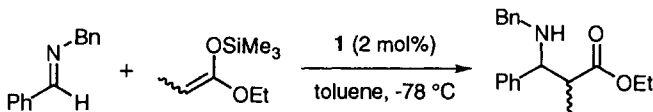
Scheme 3.3

aldehydes, and (*E*)- and (*Z*)-ketene silyl acetals give *anti* and *syn* products as major diastereomers, respectively.

Protocol 1.

Aldol-type reaction of ketene silyl acetals with *N*-benzylimines catalysed by **1**⁵ (Scheme 3.4)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.4

Equipment

- Two-necked, round-bottomed flask (25 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Three-way stopcock
- Column for flash chromatography
- Vacuum/argon source

Materials

- Anhydrous **1** (0.247 M) in toluene,^a 24 μ L, 0.006 mmol
- *N*-Benzylidenebenzylamine (FW 195.3),⁵ 116.8 mg, 0.6 mmol
- 1-Ethoxy-1-(trimethylsiloxy)propene⁵ (FW 174.3), 208.9 mg, 1.2 mmol
- Dry toluene^b
- Silica gel for flash chromatography, Merck 9385

flammable toxic
moisture sensitive
flammable
flammable toxic
irritant dust

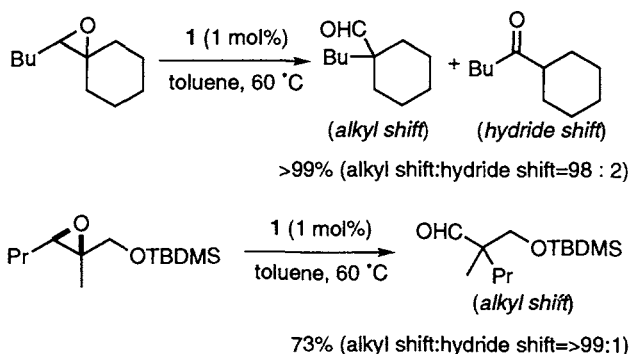
1. Add an anhydrous solution of **1** in toluene dropwise at -78°C to a solution of *N*-benzylidenebenzylamine and 1-ethoxy-1-(trimethylsiloxy)propene (*E/Z* = 85:15) in dry toluene (6 mL).
2. After stirring for 13.5 h at -78°C in a dry ice/methanol bath, warm the reaction mixture to 25°C and stir for another 2 h.
3. After pouring aqueous sodium hydrogencarbonate (0.1 mL) into the resultant solution, dry the mixture over MgSO_4 , filter, and concentrate under vacuum.
4. Purify the crude oil by column chromatography on silica gel (eluant: hexane-ethylacetate, 15:1) to produce ethyl 3-benzylamino-2-methyl-3-phenylpropanoate (178 mg, >99% yield) as a colourless oil. The product is >98% pure by ^1H NMR, IR analysis, and may be characterized further by elemental analysis. *Syn/anti* ratio of the products is 36:64 by ^1H NMR assay.

^a From Tosio-Akuzo Chemical Co. Ltd, Japan; used as received.

^b Distil toluene from calcium hydride under argon.

2.1.3 Stereoselective rearrangement of epoxides⁶

The protic or Lewis acid-promoted rearrangement of epoxides to carbonyl compounds is a well-known synthetic transformation. $\text{BF}_3 \cdot \text{OEt}_2$ appears to be the most widely used Lewis acid for rearrangement. It is often consumed or altered in the course of these reactions, and is thus a reagent rather than a catalyst, although less than an equivalent is effective in some instances. We have found that **1** is a highly efficient catalyst in the rearrangement of epoxides. The rearrangement of trisubstituted epoxides proceeds successfully in the presence of catalytic amounts of **1** (anhydrous grade) *via* a highly selective alkyl shift to give the corresponding aldehydes (Scheme 3.5). The exceptional bulkiness of **1** may be effective in the selective rearrangement of epoxides *via* an alkyl shift.



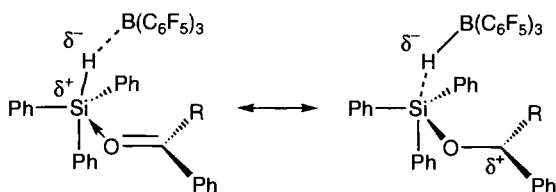
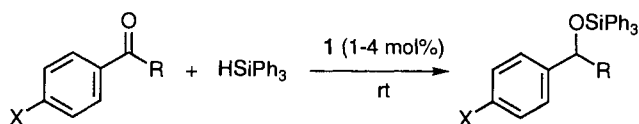
Scheme 3.5

2.1.4 Hydrosilation of aromatic aldehydes, ketones, and esters⁷

Hydrosilation of carbon–oxygen bonds is a mild method for selective reduction of carbonyl functions. Parks and Piers⁷ found that aromatic aldehydes, ketones, and esters are hydrosilylated at room temperature in the presence of 1–4 mol% of **1** and 1 equiv of Ph_3SiH . The reduction takes place by an unusual nucleophilic/electrophilic mechanism: the substrate itself serves to nucleophilically activate the Si–H bond, while hydride transfer is facilitated by **1** (Scheme 3.6).

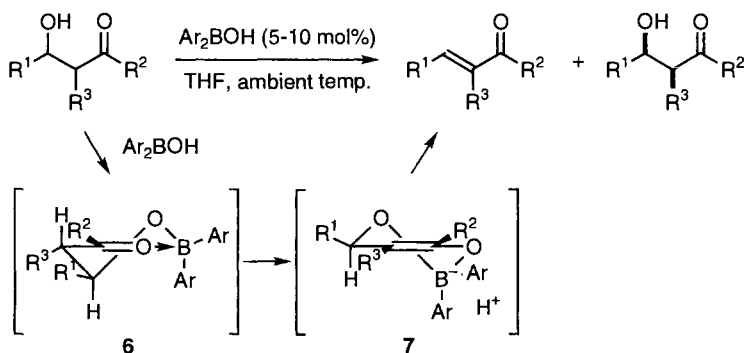
2.2 Diarylborinic acids as efficient catalysts for selective dehydration of aldols⁸

Diarylborinic acids with electron-withdrawing aromatic groups are effective for Mukaiyama aldol condensation.⁸ The catalytic activities of diarylborinic acids **2** and **3** are much higher than those of the corresponding arylboronic acids. Recently, we developed the selective dehydration of β -hydroxy carbonyl compounds catalysed by diarylborinic acids.⁸ The dehydration is strongly promoted in tetrahydrofuran (THF). In most cases, the reaction proceeds smoothly, and

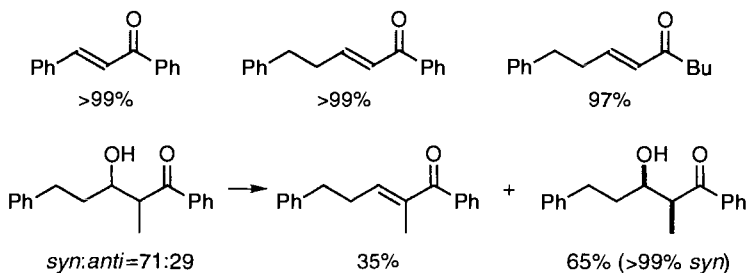


Scheme 3.6

α,β -enones are obtained as *E* isomers in high yields. In the reaction of α -substituted- β -hydroxy carbonyl compounds, α,β -enones are preferentially obtained from *anti*-aldols, and most of the *syn*-aldols are recovered. Thus, the present dehydration is a useful and convenient method for isolating pure *syn*-aldol from *syn*- and *anti*-isomeric mixtures. The transformation to α,β -enones occurs *via* an enolate intermediate **7** derived from the selective deprotonation of a pseudo-axial α -proton perpendicular to the carbonyl face of a cyclic intermediate **6** (Scheme 3.7).



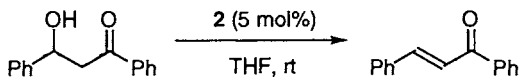
Examples



Scheme 3.7

Protocol 2.**Dehydration of aldols catalysed by 2⁸ (Scheme 3.8)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 3.8****Equipment**

- 3 Round-bottomed flask (5 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Pressure-equalized addition funnel (10 mL)
- Column for flash chromatography
- Three-way stopcock
- Vacuum/argon source

Materials

- 2^a (FW 361.9), 3.6 mg, 0.01 mmol
- 1,3-Diphenyl-3-hydroxy-1-propanone (FW 226.3), 45 mg, 0.2 mmol
- Dry THF^b
- Silica gel for flash chromatography, Merck 9385

**Toxic, white solid
irritant
flammable irritant
irritant dust**

1. Add 2 to a solution of 1,3-diphenyl-3-hydroxy-1-propanone in THF (1 mL) and stir the reaction mixture at room temperature for 2 h.
2. Pour the reaction mixture into 1M NaOH aq., and extract with CH₂Cl₂ three times. Dry the combined organic extracts over MgSO₄ filter, and concentrate in vacuum.
3. Purify the crude product by column chromatography on silica gel using a mixture of hexane and ethyl acetate (5/1) as eluent to give chalcone (40 mg, 0.19 mmol, 96%). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized further by elemental analysis.

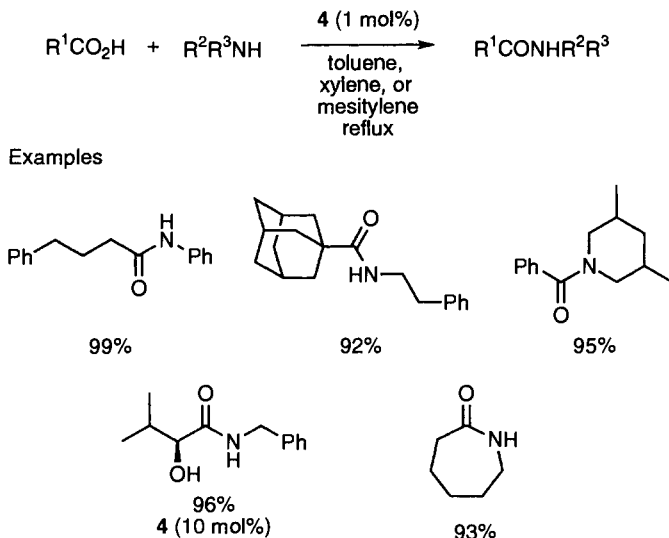
^a2 can be prepared by hydrolysis of the known bis(pentafluorophenyl)boron chloride. Chambers, R. D.; Chivers, T. J. *Chem. Soc.* **1965**, 3933.

^bDistil THF from sodium and benzophenone under argon.

2.3 3,4,5-Trifluorophenylboronic acid as an amidation catalyst⁹

There are several different routes to carboxamides. In most cases, a carboxylic acid is converted into a more reactive intermediate, e.g. acid chloride, which is then allowed to react with an amine. For practical reasons, it is preferable to form the reactive intermediate *in situ*. We have found that arylboronic acids with electron-withdrawing groups **4** and **5** act as highly efficient

catalysts in the amidation between carboxylic acids and amines (Scheme 3.9). The catalytic amidation of optically active aliphatic α -hydroxycarboxylic acids with benzylamine proceeds with no measurable loss of enantiomeric purity.

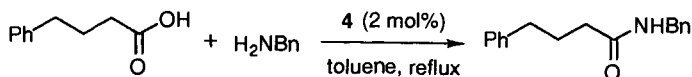


Scheme 3.9

Protocol 3.

Amidation reaction of carboxylic acids and amines catalysed by **4**⁹ (Scheme 3.10)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.10

Equipment

- Round-bottomed flask (50 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)

- Pressure-equalized addition funnel (10 mL)
- Column for flash chromatography
- Three-way stopcock
- Vacuum/argon source
- Reflux condenser
- Oil bath

Materials

- **4**^a (FW 175.9), 8.8 mg, 0.05 mmol
- 4-Phenylbutyric acid (FW 164.2), 821 mg, 5 mmol
- Benzylamine (FW 107.16), 546 μ L, 5 mmol
- Dry toluene^a
- 4-Å molecular sieves ca. 4 g

toxic, white solid
irritant
flammable
flammable toxic
pellets

1. Add 4-phenylbutyric acid, benzylamine, and **4** in 25 mL of toluene to a dry, round-bottomed flask fitted with a stirrer bar and a pressure-equalized addition funnel (containing a cotton plug and 4-Å molecular sieves and functioning as a Soxhlet extractor) surmounted by a reflux condenser.
2. Bring the mixture to reflux by the removal of water.
3. After 18 h, cool the resulting mixture to ambient temperature, wash with aqueous ammonium chloride and aqueous sodium hydrogencarbonate successively, and extract the product with ethyl acetate. Dry the combined organic layers over magnesium sulfate. Evaporate the solvent, and purify the residue by column chromatography on silica gel (eluant: hexane–ethyl acetate = 3:1) to give *N*-benzyl-4-phenylbutyramide (1.221 g, 96% yield). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized further by elemental analysis.

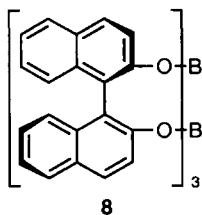
^a Distil toluene from calcium hydride under argon.

3. Chiral boron reagents as Lewis acids

3.1 Catalytic enantioselective carbo-Diels–Alder reactions

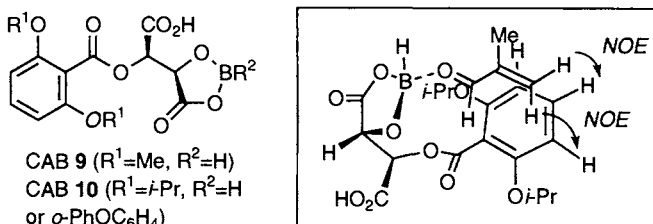
The asymmetric Diels–Alder reaction is now of great interest because of its potential to introduce some asymmetric centres simultaneously during carbon–carbon bond formation.

Kaufmann and Boese¹⁰ developed asymmetric catalyst **8** derived from H₂BBr–SMe₂ and 1,1-binaphthol. The diborate structure with a propeller-like shape has been established by X-ray analysis. The reaction between methacrolein and cyclopentadiene catalysed by **8** gives the cycloadducts with 97% exo selectivity and 90% ee.



Using another promising approach^{11–16} it was found that an (acyloxy) borane RCO₂BR'₂ behaves as a Lewis acid, and the chiral (acyloxy)borane

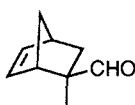
(CAB) **9** is an excellent asymmetric catalyst for the Diels–Alder reaction between cyclopentadiene and acrylic acid¹¹ or methacrolein^{12,13} (Scheme 3.11). The α -substituent on the dienophile increases the enantioselectivity. When there is a β -substitution on the dienophile, the cycloadduct is almost racemic, but for a substrate having substituents at both α - and β -positions, high ee's have been observed. According to NOE studies, the effective shielding of the *si*-face of the co-ordinated α,β -enal arises from π -stacking of the 2,6-di-isopropoxybenzene ring and the co-ordinated aldehyde.



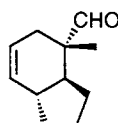
Examples



9 (10 mol%)
exo/endo: 4/96
endo: 78% ee



9 (10 mol%)
exo/endo: 89/11
exo: 96% ee



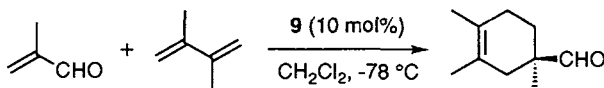
9 (10 mol%)
exo/endo: 4/96
exo: 92% ee

Scheme 3.11

Protocol 4.

Enantioselective Diels–Alder reaction catalysed by CAB **9**¹³ (Scheme 3.12)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.12

Equipment

- Three-necked round-bottomed flask (100 mL)
- Separating funnel (500 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (500 mL)
- Pressure-equalized addition funnel (10 mL)
- Column for flash chromatography
- Three-way stopcock
- Vacuum/argon source
- Rubber septum

Materials

- Mono (2,6-dimethoxybenzoyl) tartaric acid¹³ (FW 314.2),
1.57 g, 5 mmol **irritant**
- Borane–THF (1.40 M)^a, 3.57 mL, 5 mmol **flammable liquid, moisture-sensitive**
- Methacrolein (FW 70.09)^b, 4.14 mL, 50 mmol **flammable liquid, corrosive**
- 2,3-Dimethyl-1,3-butadiene (FW 82.15)^c, 8.49 mL, 75 mmol **flammable liquid**
- Dry dichloromethane^d **flammable toxic**
- 4-Å molecular sieves ca. 4 g **pellets**

1. Equip a three-necked round-bottomed flask containing a magnetic stirrer bar with a rubber septum and three-way stopcock with an argon inlet. Repeat flushing with dry argon to displace the air.
2. Charge the flask with mono(2,6-dimethoxybenzoyl) tartaric acid and 50 mL of dry dichloromethane, and cool in an ice bath.
3. Through the septum, with a syringe, add dropwise a borane–THF solution at 0°C over a period of 30 min.
4. Stir the reaction mixture for 15 min at 0°C and then cool to –78°C in a dry ice/methanol bath.
5. Add freshly distilled methacrolein to this solution via a syringe dropwise.
6. After the addition is complete, introduce 2,3-dimethyl-1,3-butadiene to the solution at the same temperature and stir the mixture for 12 h.
7. Pour the cold reaction mixture into 150 mL of ice-cold saturated sodium bicarbonate and extract the product with three 200-mL portions of hexane. Wash the combined organic phase with brine (2 × 200 mL), dry over sodium sulfate, filter, and concentrate at atmospheric pressure. Distil the residue at reduced pressure to afford (1*R*)-1,3,4-trimethyl-3-cyclohexene-1-carboxaldehyde (6.53 g, 86%) as a colourless liquid, b.p. 92–93°C (23 mm). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized further by elemental analysis.^e

^a Titrate borane–THF complex which can be obtained from Tosoh-Akzo Chemical Company, Ltd. in Japan before use. Vigorous evolution of hydrogen is observed during addition of borane–THF solution to the reaction mixture.

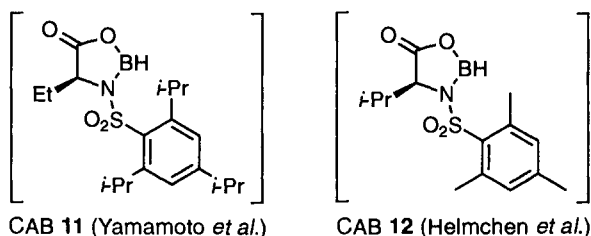
^b Dry methacrolein from Tokyo Kasei Kogyo Company, Ltd with calcium sulfate and distil through a 20-cm Vigreux column under argon prior to use.

^c Distil 2,3-dimethyl-1,3-butadiene, from Tokyo Kasei Kogyo Company Ltd, before use.

^d Distil dichloromethane from calcium hydride under argon.

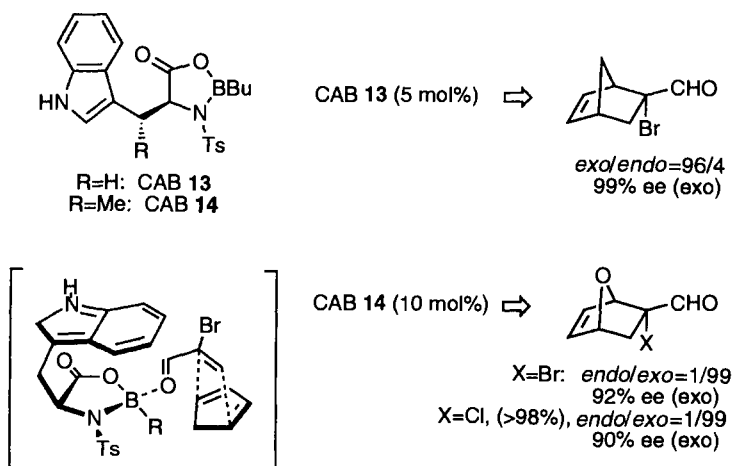
^e The optical purity of this adduct is 95% as determined by 200 MHz ¹H NMR spectroscopy and GC analysis (capillary column PEG, 0.25 mm × 25 m, purchased from Gaskuro Kogyo Company, Ltd in Japan) after conversion into the corresponding chiral acetal as follows. Stir a solution of the adduct, (2*R*,4*R*)-2,4-pentanediol (1.2 equiv, obtained from Wako Pure Chemical Industries), triethyl orthoformate (1.2 equiv.), and *p*-toluenesulfonic acid monohydrate (as a 5 mM solution) in dry benzene at ambient temperature for 3 h. Pour the mixture into saturated sodium bicarbonate and extract the product with ether. Dry the combined organic phases over sodium sulfate and concentrate on a rotary evaporator. Purify the residue by flash column chromatography on silica gel using hexane–ethyl acetate (25:1) as eluant to give the acetal quantitatively.

In 1991, Helmchen *et al.*^{17,19} and our group¹⁸ found that *N*-sulfonyl derivatives of α -amino acid react with diborane, giving complexes formulated as CAB (Scheme 3.13). These CAB complexes catalyse asymmetric cycloadditions between α,β -enals and dienes.



Scheme 3.13

A similar effect was also published by Corey *et al.* on CAB **13**.^{20–24} Especially efficient is the asymmetric catalysis of cycloaddition between 2-bromoacrolein and various dienes (>90–95% ee). The transition-state is believed to be as represented in Scheme 3.14.²¹ Attractive interactions between the indolyl moiety and the π -acidic dienophile protect one face of the dienophile. CAB **14** derived from *N*-tosyl (αS , βR)- β -methyltryptophan catalyses the cycloaddition of 2-bromoacrolein and furan with 92% ee.²⁴

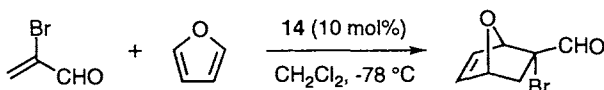


Scheme 3.14

Protocol 5.

Enantioselective Diels–Alder reaction catalysed by CAB 14²⁴
(Scheme 3.15)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.15

Equipment

- Three-necked round-bottomed flask
- Separating funnel
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks
- Column for flash chromatography
- Three-way stopcock
- Soxhlet extractor
- Reflux condenser
- Rubber septum
- Vacuum/nitrogen source

Materials

- *N*-Tosyl (α S, β R)-methyltryptophan²¹ (FW 358.4), 2.06 g, 5.6 mmol
- Butylboronic acid (FW 101.9), 0.60 g, 6.7 mmol
- 2-Bromoacrolein (FW 135.0)²⁰
- Furan (FW 68.1)
- Dry toluene^a
- Dry THF^b
- Dry dichloromethane^a
- CaH₂
- Sand

irritant
hygroscopic
flammable liquid, corrosive
highly toxic cancer suspect agent
flammable liquid, toxic
flammable liquid, irritant
toxic, irritant
flammable solid, moisture sensitive

1. Add *N*-tosyl (α S, β R)-methyltryptophan and butylboronic acid in a round-bottomed flask equipped with magnetic stirrer and a Soxhlet extractor with reflux condenser with a rubber stopper at the top, connect to a vacuum manifold and place under nitrogen.
2. A thimble in the extractor contains layers of sand (3 cm, bottom) and CaH₂ (3 cm, top).
3. After the addition of 20 mL of dry THF and 40 mL of dry toluene, heat the mixture at rapid reflux for 20 h using an oil bath heated to >165 °C.
4. Quickly disconnect the reaction flask and attach to a vacuum line, and remove the solvent *in vacuo* leaving catalyst **14** as a colourless viscous oil. Store **14** in a tightly sealed flask as a 0.1 M solution in 9:1 toluene–THF.
5. Remove the solvent *in vacuo* and exchange with CH₂Cl₂ toluene or other solvent just prior to the Diels–Alder reaction. THF serves to stabilize toluene solutions of catalyst **14**.

Protocol 5. Continued

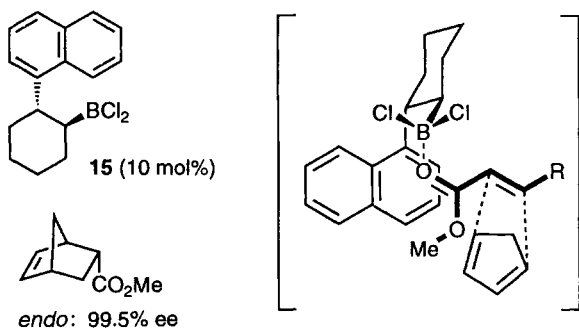
6. The reaction of 5 equiv. of furan with 2-bromoacrolein in the presence of 10 mol% of **14** in dichloromethane at -78°C is complete in 5 h and gives the Diels–Alder adduct in 98% yield and 96:4 enantioselectivity, as determined by 500 MHz analysis of the α -methoxy- α -(trifluoromethyl)phenylacetic ester of the corresponding primary alcohol (from NaBH_4 reduction of the aldehyde). The *N*-tosyl carboxylic acid precursor of **14** is efficiently recovered for reuse in each case.

^a Distil toluene and dichloromethane from calcium hydride under nitrogen.

^b Distil THF from sodium and benzophenone under nitrogen.

Itsuno *et al.* explored the possibility of using polymer-supported chiral Lewis acids in the cycloaddition of methacrolein with cyclopentadiene.^{25–28} Using insoluble polymer-supported CAB, high enantioselectivity (up to 95% ee) has been achieved.²⁸

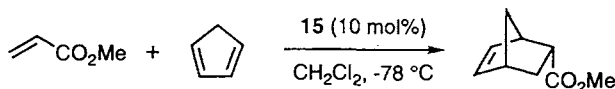
Hawkins *et al.*^{29,30} developed a simple and efficient catalyst for Diels–Alder reaction based on a chiral alkyldichloroborane **15** (Scheme 3.16). They have predicted the approach of the diene on one of the faces of the methyl crotonate because the other face is protected by π – π donor-acceptor interactions on the basis of the crystal structure study of a molecular complex between methyl crotonate and **15**.



Scheme 3.16

Protocol 6.**Enantioselective Diels–Alder reaction catalysed by chiral alkyldichloroborane 15³⁰ (Scheme 3.17)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 3.17****Equipment**

- Schlenk flask (100 mL)
- Separating funnel (100 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (100 mL)
- Column for flash chromatography
- Vacuum/nitrogen source

Materials

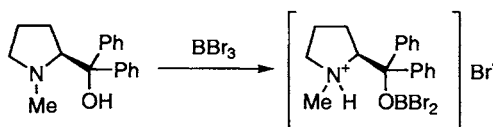
- (1*R*,2*R*)-**15**^{29,30} (FW 291.0), 352.1 mg, 1.21 mmol
- Cyclopentadiene (FW 66.1), 4.66 g, 70.5 mmol
- Methyl acrylate (FW 86.1), 1.21 g, 14.1 mmol
- Dry dichloromethane

moisture sensitive
flammable liquid, toxic
flammable liquid, lachrymator
toxic, irritant

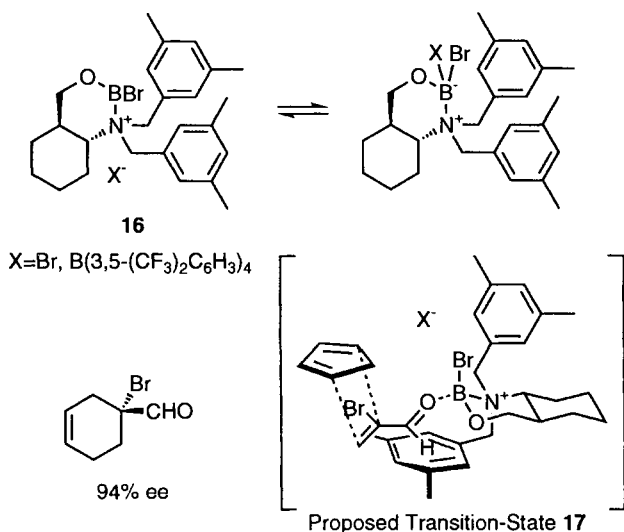
1. Add (1*R*,2*R*)-**15** in a Schlenk flask, cool the flask to -78°C , and add 14 mL of freshly distilled CH_2Cl_2 slowly. Add enough CH_2Cl_2 to establish a solvent to diene ratio of 3:1 (v/v).
2. Then, add cyclopentadiene and methyl acrylate successively.
3. Seal the reaction flask and allow to warm to the reaction temperature.
4. Quench the reaction with 10% NaHCO_3 (aq) and isolate the products (95% yield) by flash chromatography. Determine enantiomeric excess with chiral GC (97% ee).^a

^a Chiral GC analyses are performed with J&W Cyclodex-B β -cyclodextrin column.

Mukaiyama *et al.* have found that prolinol derivatives combined with BBr_3 are good chiral catalysts for some Diels–Alder reactions.^{31,32} For example, methacrolein and cyclopentadiene afford the *exo* adduct (*exo:endo* = >99:1) in 97% ee (20 mol% of catalyst). The chiral catalyst is believed to be the HBr adduct salt of the amino boron derivative (Scheme 3.18).

**Scheme 3.18**

Recently, a new chiral Lewis acid **16** was developed, and its utility was demonstrated in cycloadditions with both reactive and unreactive 1,3-dienes (Scheme 3.19).³³ In the hypothetical transition-state model **17**, one of the *N*-CH₂Ar substituents serves to block attack on the lower face of the *s-trans*-co-ordinated dienophile whereas the other screens another region in space and limits the rotational position of dienophile and *N*-CH₂Ar moieties.

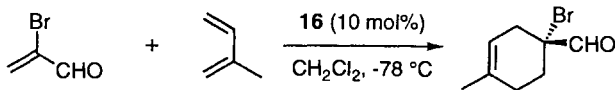


Scheme 3.19

Protocol 7.

Enantioselective Diels–Alder reaction catalysed by chiral cationic Lewis acid **16**³³ (Scheme 3.20)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.20

Equipment

- Vacuum/argon source

- There is no information about apparatus.³³

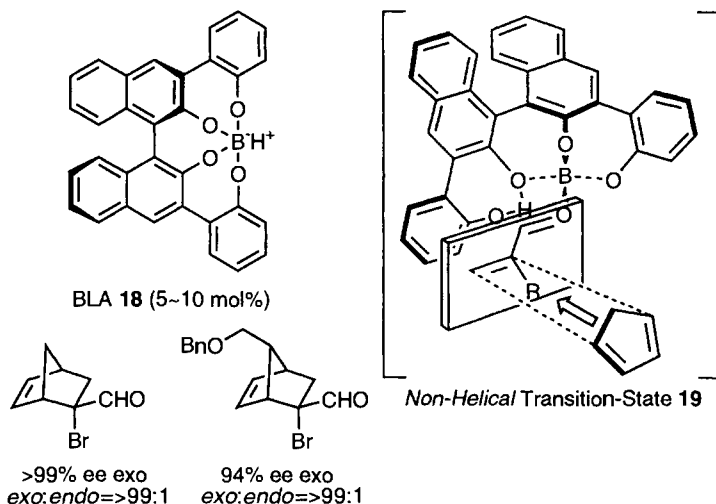
Materials

- The silyl ether of chiral ligand (FW 423.7),³³ 40.3 mg, 0.095 mmol
 - Boron tribromide (FW 250.5), 0.46 M in CH₂Cl₂ 187 μ L, 0.086 mmol
 - Ag⁺B[C₆H₃-3,5-(CF₃)₂]₄⁻ (FW 971.1),^a 83.5 mg, 0.086 mmol
 - 2-Bromoacrolein (FW 135.0), 76 μ L, 0.94 mmol
 - Isoprene (FW 68.1), 500 μ L 5 mmol
 - Dry dichloromethane
- corrosive moisture-sensitive
toxic
flammable liquid, irritant
cancer suspect agent, flammable liquid
toxic, irritant

1. Add a solution of BBr₃ in CH₂Cl₂ to a solution of the silyl ether of chiral ligand in 1 mL of CH₂Cl₂ at -94 °C (hexane-liquid N₂ bath) under dry argon, and after 5 min, replace the cooling bath with dry ice-acetone to bring the temperature to -78 °C.
2. After adding a freshly prepared dry solution of Ag⁺B[C₆H₃-3,5-(CF₃)₂]₄⁻ in 1 mL of CH₂Cl₂, stir the reaction mixture for 20 min at -78 °C and then cool to -94 °C.
3. Add 2-bromoacrolein and isoprene successively (each dropwise), and stir the reaction mixture for 1 h at -94 °C and then quench with 150 μ L of triethylamine.
4. After warming the reaction mixture to room temperature and removing the inorganic salts by filtration, evaporate the solvents, and purify the residue by chromatography on silica gel to give 191 mg (99%) of the isoprene Diels-Alder adduct, [α]_D²³ + 82.2° (c 0.8, CH₂Cl₂) and also recover ligand (28.6 mg, 83%).

^a Prepare the silver salt as follows. Shake an ethereal solution of NAB[C₆H₃-3,5-(CF₃)₂]₄ (Broolhart, M.; Grant, B.; Volpe, A. F., Jr *Organometallics* **1992**, *11*, 3920) with 2 equiv. of aqueous AgNO₃ in a separating funnel for 5 min, and separate the layers. Evaporation of the ether layer affords a quantitative yield of the colourless silver salt which is dissolved in ether to give a clear 0.1 M solution. Store this at -78 °C in a flask wrapped with aluminium foil to exclude light. To prepare the catalyst, concentrate a measured amount of this ethereal solution *in vacuo*, dissolve in dry CH₂Cl₂, and dry over activated molecular sieves 4 Å for 1 h at room temperature (with constant protection from light).

We found that Brønsted acid-assisted chiral Lewis acid (BLA) **18** achieved high selectivity through the double effect of intramolecular hydrogen binding interaction and attractive π - π donor-acceptor interaction in the transition-state (Scheme 3.21).³⁴ Extremely high enantioselectivity (>99-92% ee) and *exo*-selectivity (>99-97% *exo*) are obtained for cycloadditions of α -substituted α,β -enals with dienes. The absolute stereopreference in the reaction can be easily understood in terms of the most favourable transition-state assembly **19**. The co-ordination of a proton of 2-hydroxyphenyl group with an oxygen of the adjacent B-O bond in complex **19** plays an important role in asymmetric induction; this hydrogen binding interaction via Brønsted acid causes Lewis acidity of boron and π -basicity of phenoxy moiety to increase.

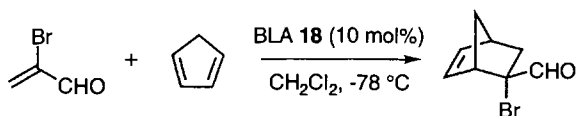


Scheme 3.21

Protocol 8.

Enantioselective Diels–Alder reaction catalysed by chiral BLA **18**³⁴ (Scheme 3.22)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.22

Equipment

- Round-bottomed flask (10 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Column for flash chromatography
- Vacuum/argon source
- Pressure-equalized addition funnel
- Three-way stopcock

Materials

- (*R*)-3,3'-bis(2-hydroxyphenyl)-2,2'-dihydroxy-1,1'-binaphthyl (FW 472.5),³⁴ 23.5 mg, 0.05 mmol
- Trimethyl borate (FW 103.9), 0.1 M in CH_2Cl_2 0.5 mL, 0.05 mmol
- 2-bromoacrolein (FW 135.0), 80.8 μL , 1 mmol
- Cyclopentadiene (FW 66.1), 332 μL , 4 mmol

white solid, irritant

flammable liquid, moisture-sensitive
flammable liquid, irritant
flammable liquid, toxic

3: Boron reagents

- Dry dichloromethane^a
- Dry THF^b
- 4-Å molecular sieves

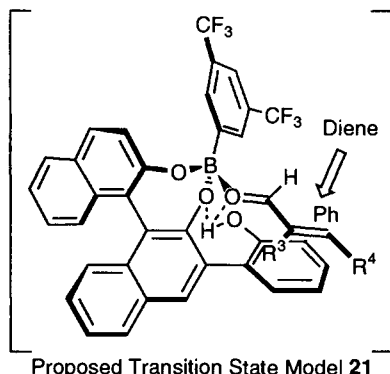
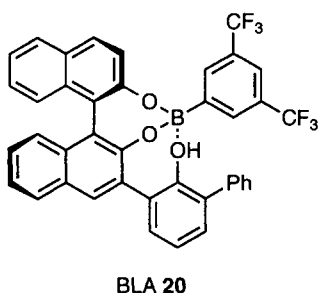
toxic, irritant
flammable irritant
pellets

1. Add (*R*)-3,3'-bis(2-hydroxyphenyl)-2,2'-dihydroxy-1,1'-binaphthyl, trimethyl borate, and 3 mL of dichloromethane in a dry round-bottomed flask fitted with a stirrer bar and a pressure-equalized addition funnel (containing a cotton plug and ca. 4 g of 4-Å molecular sieves (pellets) and functioning as a Soxhlet extractor) surmounted by a reflux condenser.
2. Secure an argon atmosphere, and bring the solution to reflux (bath temperature 50–60 °C).
3. After 2 h, cool the reaction mixture to 25 °C and quickly remove the addition funnel and condenser and replace with a septum.
4. Add dry THF to the white precipitate in dichloromethane at 25 °C, and after 2 h the precipitate is completely dissolved.
5. After cooling a colourless solution of the catalyst (*R*)-**18** to –78 °C, add dropwise 2-bromoacrolein and cyclopentadiene.
6. After 4 h, add 50 mL of H₂O and warm the mixture to 25 °C, dry over MgSO₄, filter, and purify by eluting with hexane/ethyl acetate (10:1) to afford 201 mg of Diels–Alder adduct (1*S*, 2*S*, 4*S*)-bromo aldehyde as a white solid (1.0 mmol, >99% yield, exo:endo = >99:1, >99% ee) and quantitative recovery of pure chiral ligand. Determine enantioselectivity by reduction with NaBH₄, conversion to the Mosher ester, and ¹H NMR and HPLC analysis (Daicel AD).

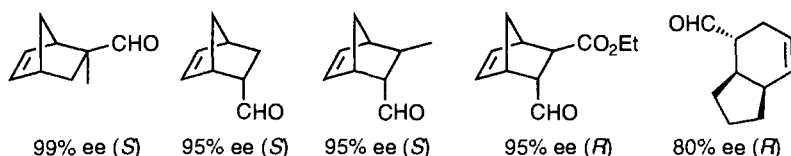
^a Distil dichloromethane from calcium hydride under argon.

^b Distil THF from sodium and benzophenone under argon.

Diels–Alder reactions of α -unsubstituted α,β -enals with BLA **18** as well as most chiral Lewis acids exhibit low enantioselectivity and/or reactivity. We have developed a new type of BLA, **20** which was prepared from a chiral triol and **5** (Scheme 3.23).³⁵ **20** is extremely effective in enantioselective cycloaddition of both α -substituted and α -unsubstituted α,β -enals with various dienes. The Brønsted acid in the BLA clearly accelerates the cycloaddition. The high enantioselectivity and stereochemical results attained in this reaction can be understood in terms of the transition-state model **21**.



Examples

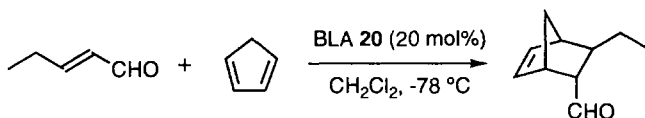


Scheme 3.23

Protocol 9.

Enantioselective Diels–Alder reaction catalysed by BLA 20³⁵
(Scheme 3.24)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.24

Equipment

- Schlenk flasks (10 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Column for flash chromatography
- Vacuum/argon source
- Pressure-equalized addition funnel
- Three-way stopcock
- Oil bath

Materials

- The chiral ligand (FW 454.5),³⁵ 54.5 mg, 0.12 mmol **white solid**
- Monomeric **5** (0.043 M) in CH₂Cl₂-THF-H₂O (20:3:0.054), 23.3 μ L, 0.1 mmol **hygroscopic**
- (*E*)-2-Pentenal (FW 84.12), 42.1 mg, 0.5 mmol **flammable liquid, irritant**
- Cyclopentadiene (FW 66.1), 166 μ L, 2 mmol **flammable liquid, toxic**
- Dry THF^a **flammable irritant**
- Dry dichloromethane^b **toxic, irritant**
- 4-Å molecular sieves **powder**

1. Stir a mixture of the chiral ligand and a solution of monomeric **5** in CH₂Cl₂-THF-H₂O (20:3:0.054) at ambient temperature for 2 h.
2. Transfer the resulting colourless solution into a Schlenk tube containing c. 0.5 mL of anhydrous dichloromethane and MS 4Å (powder, 250 mg, activated by heating at 200°C under vacuum [c. 3 Torr] for 12 h), and use a further c. 0.5 mL of dichloromethane to transfer the residue into the Schlenk.
3. Stir the mixture at ambient temperature for another 12 h.
4. Then, evaporate the solvents and heat the resulting solid to 100°C (oil bath) for 2 h under vacuum (c. 3 Torr) to dry catalyst.
5. After cooling to ambient temperature, purge the flask with argon and add dichloromethane (2 mL).
6. Cool the mixture to -78°C, add dropwise (*E*)-pentenal, and 1 min later add freshly distilled cyclopentadiene slowly along the wall of the flask.
7. After stirring the reaction mixture at -78°C for 72 h, quench the reaction with pyridine (20 μ L, 0.25 mmol), warm to ambient temperature, and filter to remove molecular sieves. Wash the filtrate with ether, dry over MgSO₄, and concentrate to afford the crude products. Purify by silica gel chromatography eluting with pentane-ether to provide the pure Diels-Alder adduct (73% yield, *endo/exo* = 91:9, 98% ee for *endo*-isomer). Determine the *exo/endo* ratio by ¹H NMR analysis (500 MHz): δ 9.38 (d, *J* = 3.3 Hz, 1H, CHO (*endo*)), 9.79 (d, *J* = 2.9 Hz, 1H, CHO (*exo*)). Determine the ee by acetalization with (-)-(2*R*,4*R*)-2,4-pentanediol and GC analysis (90°C, PEG-HT Bonded (25 m \times 0.25 mm)): *t*_R = 29.0 min (major *endo*-isomer), 36.0 min (minor *endo*-isomer), 37.8 min (minor *exo*-isomer), 38.5 min (major *exo*-isomer). The absolute configuration is not established.

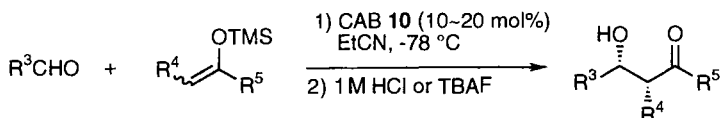
^a Distil YHF from sodium and benzophenone under argon.^b Distil dichloromethane from calcium hydride under argon.

3.2 Catalytic enantioselective aldol and allylation reactions

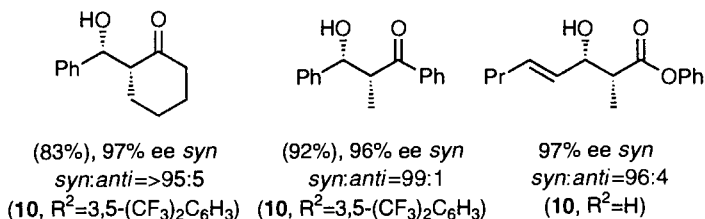
Asymmetric aldol and allylation reactions are now of great interest because of their utility for introduction of asymmetric centres and functional groups.

We have reported that CAB **10**, R² = H, is an excellent catalyst (20 mol%)

for the enantioselective and diastereoselective Mukaiyama condensation of simple enol silyl ethers with various aldehydes.^{36,37} The reaction is accelerated without reducing the enantioselectivity by using 10–20 mol% of **10**, $R^2 = 3,5\text{-(CF}_3)_2\text{C}_6\text{H}_3$. The enantioselectivity is increased without reducing the chemical yield by using 20 mol% of **10**, $R^2 = o\text{-PhOC}_6\text{H}_4$. Another aldol-type reaction of ketene silyl acetals derived from phenyl esters with achiral aldehydes also proceeds smoothly with **10** and can furnish *syn* β -hydroxy esters with high optical purity.^{37,38} Regardless of the stereochemistry of enol silyl ethers, *syn* aldols are highly selectively obtained *via* the acyclic extended transition-state mechanism (Scheme 3.25).



Examples

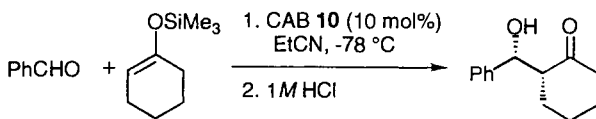


Scheme 3.25

Protocol 10.

Enantioselective Mukaiyama Aldol reaction catalysed by CAB **10**³⁷ (Scheme 3.26)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.26

Equipment

- Schlenk flask (10 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Column for flash chromatography
- Vacuum/argon source
- Three-way stopcock
- Rubber septum

Materials

- Mono(2,6-diisopropoxybenzoyl)tartrac acid¹³
(FW 370.4), 74.1 mg, 0.2 mmol
- **5** (FW 257.9), 51.6 mg, 0.2 mmol
- 1-Trimethylsiloxy-1-cyclohexene (FW 170.3),
204.4 mg, 1.2 mmol
- Benzaldehyde (FW 106.12), 106.1 mg, 1.0 mmol
- Dry propionitrile^a

irritant
hygroscopic

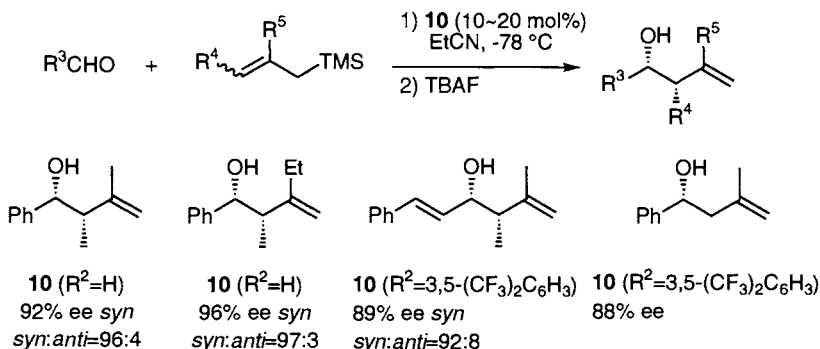
flammable liquid

highly toxic cancer suspect agent
highly toxic flammable liquid

1. Equip a Schlenk flask containing a magnetic stirrer bar with a rubber septum and three-way stopcock with an argon inlet. Repeat flushing with dry argon to displace the air.
2. Add mono(2,6-di-isopropoxybenzoyl)tartrac acid and **5** in the flask, dissolve in 1 mL of dry propionitrile, stir the resulting solution at 25°C for 30 min, and cool the reaction flask to -78°C.
3. Add 1-trimethylsiloxy-1-cyclohexene and benzaldehyde successively and stir the reaction mixture for 12 h at low temperature.
4. Pour this cold solution into water and extract the product with ether repeatedly. Dry the combined ether layers over MgSO₄, concentrate *in vacuo* and treat the residue with 1 M HCl (aq)-THF solution (2 mL, 1/1 in vol.). After usual work-up, purify the crude product by column chromatography on silica gel to give aldol adducts (83% yield, *syn:anti* = >95:5, 97% ee for *syn* diastereomer). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized further by elemental analysis. Determine the *syn/anti* ratio by ¹H NMR analysis (500 MHz), and determine the ee by ¹H NMR analysis of (+)-MTPA ester.

^a Distil propionitrile from calcium hydride under argon.

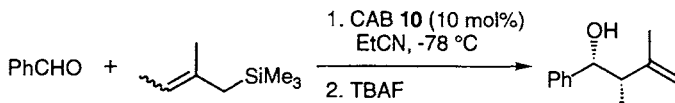
We have found that CAB **10** has a powerful activity for the Sakurai-Hosomi allylation reaction of aldehydes to furnish homoallylic alcohols in excellent enantiomeric excess (Scheme 3.27).³⁹ γ -Alkylated allylsilanes exhibit excellent diastereo- and enantioselectivities affording *syn* homoallylic alcohols of higher optical purity. Regardless of the geometry of starting allylsilanes, the predominant isomer in this reaction has *syn* configuration. The observed preference for relative and absolute configurations for the adducts is predicted on the basis of an extended transition-state model similar to that for the CAB **10** catalysed aldol reaction.³⁶⁻³⁸ The 3,5-bis(trifluoromethyl)phenyl group is the most effective *B*-substituent of **10**.⁴⁰



Scheme 3.27

Protocol 11.**Enantioselective Sakurai–Hosomi allylation reaction catalysed by CAB 10⁴⁰ (Scheme 3.28)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.28

Equipment

- Schlenk flasks (10 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Column for flash chromatography
- Vacuum/argon source
- Three-way stopcock
- Rubber septum

Materials

- Mono(2,6-diisopropoxybenzoyl)tartric acid¹³ (FW 370.4) 74.1 mg, 0.2 mmol
- **5** (FW 257.9), 51.6 mg, 0.2 mmol
- 1-(Trimethylsilyl)-2-methyl-2-butene (FW 142.3), 170.8 mg, 1.2 mmol
- Benzaldehyde (FW 106.12), 106.1 mg, 1.0 mmol
- Dry propionitrile^a

irritant
hygroscopic

flammable liquid
highly toxic cancer suspect agent
highly toxic flammable liquid

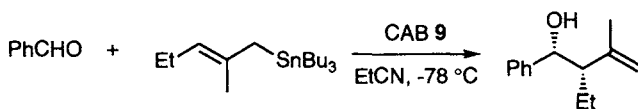
1. Equip a Schlenk flask containing a magnetic stirrer bar with a rubber septum and three-way stopcock with an argon inlet. Repeat flushing with dry argon to displace the air.

3: Boron reagents

2. Add mono(2,6-di-isopropoxybenzoyl) tartaric acid and **5** in the flask, dissolve in 1 mL of dry propionitrile, stir the resulting solution at 25°C for 30 min, and cool the reaction flask to -78°C.
3. Add 1-(trimethylsilyl)-2-methyl-2-butene and benzaldehyde successively and stir the reaction mixture for 12 h at low temperature.
4. Pour this cold solution into brine and extract with ether repeatedly, dry the combined ether layers over MgSO₄, and concentrate *in vacuo*.
5. Treat the residue with tetrabutylammonium fluoride (1.5 mL of 1 M solution in THF, 1.5 mmol) in THF (3 mL).
6. Usual work-up followed by chromatographic separation gives products (82% yield, *syn:anti* = 94.6, 91% ee for *syn* diastereomer). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized further by elemental analysis. Determine the *syn/anti* ratio and the ee by ¹H NMR analysis of (+)-MTPA ester: ¹H NMR δ 5.80 (minor enantiomer of *anti* diastereomer), 5.87 (major enantiomer of *anti* diastereomer), 5.88 (minor enantiomer of *syn* diastereomer), 5.93 (major enantiomer of *syn* diastereomer).

^a Distil propionitrile from calcium hydride under argon.

Marshall *et al.*⁴¹ reported that more reactive allyltin analogues can be used in place of allylsilane nucleophiles in our CAB **9** catalyst system, and that trifluoroacetic anhydride is an efficient promoter (Scheme 3.29).

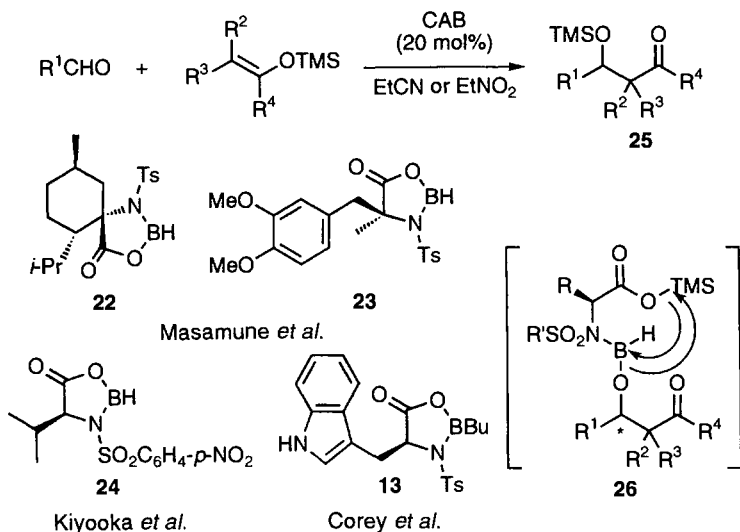


9 (20 mol%) + (CF₃)₂CO (40 mol%) : (88%), *syn:anti*=85:15, 74% ee *syn*
9 (100 mol%) + (CF₃CO)₂O (200 mol%) : (99%), *syn:anti*=90:10, 85% ee *syn*

Scheme 3.29

After the enantioselective aldol reaction using chiral oxaborolidines (another type of CAB) under a stoichiometric condition was reported by Kiyooka *et al.*⁴² Masamune *et al.*,^{43,44} Kiyooka *et al.*,⁴⁵ and Corey *et al.*⁴⁶ independently developed CAB-catalysed-aldol reactions (Scheme 3.30). Masamune *et al.* suggested that the initial aldol adduct must undergo ring closure as indicated by the arrows in **26** to release the final product **25** and to regenerate the catalyst **22** or **23**.^{43,44} In many cases, slow addition of the aldehyde to the reaction mixture has proved beneficial (permitting enough time for **26** to undergo ring closure) in improving the enantioselectivity of the reaction. Kiyooka *et al.*⁴⁵ have reported straightforward improvement of this

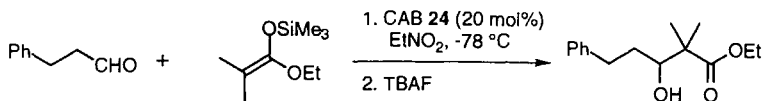
reaction to its catalytic version by employing nitroethane instead of dichloromethane as solvent.



Protocol 12.

Enantioselective Mukaiyama aldol reaction catalysed by CAB 24⁴⁵ (Scheme 3.31)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 3.31

Equipment

- Vacuum/argon source

Materials

- The sulfonamide ligand⁴⁰ (FW 302.3), 66.5 mg, 0.22 mmol
- Borane-THF (1.0 M solution in THF), 0.2 mL, 0.2 mmol
- Hydrocinnamaldehyde (FW 134.18), 134.2 mg, 1 mmol
- 1-Ethoxy-2-methyl-1-trimethylsiloxypropene (FW 188.3), 207.2 mg, 1.1 mmol
- Nitroethane

irritant
flammable liquid, moisture-sensitive
irritant

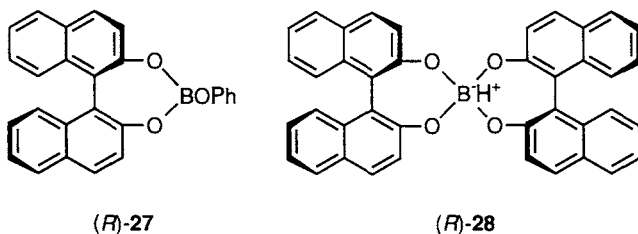
flammable liquid
flammable liquid irritant

1. Add $\text{BH}_3\text{-THF}$ complex in dropwise fashion to a solution of the sulfonamide in nitroethane at 0°C under Ar.
2. Stir the mixture for 0.5 h, allow to warm to room temperature, and stir for another 0.5 h.
3. Add a solution of hydrocinnamaldehyde in THF and a solution of the ketene silyl acetal in THF to the resulting mixture successively at -78°C .
4. Stir the mixture at -78°C for 4 h, whereupon quench the reaction by the introduction of buffer solution (pH 6.8; 5 mL). Extract the mixture with Et_2O (30 mL) twice. After the usual work-up, isolate the pure products of the silylated aldol, the corresponding aldol, and the sulfonamide by silica-gel chromatography. After desilylation with Bu_4NF (a 1.0 M solution in THF), the aldol product is obtained in 97% yield with 95% ee. Determine the ee by HPLC analysis with a chiral Daicel OD column.

Harada *et al.* reported that CAB **14** serves as an excellent catalyst for enantioselective ring-cleavage of 2-substituted 1,3-dioxolanes with silyl enol ethers.^{47–49}

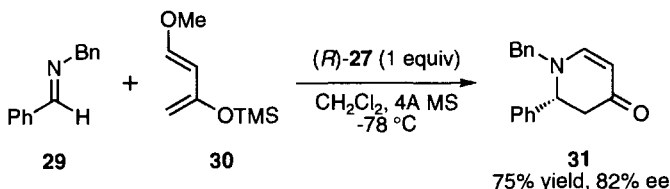
3.3 Asymmetric reactions of imines mediated by chiral boron reagents

Imines are important starting materials for nitrogen-containing compounds such as alkaloids, amino acids, β -lactams, amino sugars, and terpenes. We developed chiral Lewis acids **27** and **28** for asymmetric reactions of imines.



3.3.1 Enantioselective Diels–Alder reaction of imines mediated by chiral boron reagent **27**^{50,51}

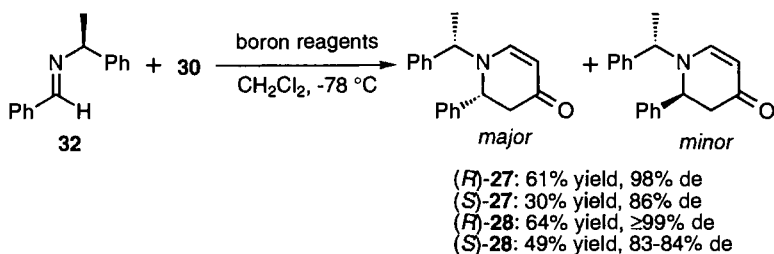
Asymmetric aza Diels–Alder reaction of imines is promoted by an *in situ* generated boron complex **27**. (*R*)-**27** is prepared *in situ* simply by mixing a 1:1 molar ratio of optically active binaphthol and triphenyl borate in dichloromethane. The aza Diels–Alder reaction of imine **29** with Danishefsky diene **30** is promoted by (*R*)-**27** at -78°C to give the adduct **31** in 75% yield with 82% ee (Scheme 3.32). The reaction is applicable to aromatic and aliphatic imines and Danishefsky-type dienes.



Scheme 3.32

3.3.2 Double asymmetric induction of Diels–Alder reaction of imines mediated by chiral boron reagents^{50–53}

There have been a number of investigations into reactions exhibiting diastereofacial selectivity with imines containing a chiral auxiliary. We succeeded in the double asymmetric induction of the aza Diels–Alder reaction of chiral imines derived from α -methylbenzylamine mediated by **27**, and thereafter found that new chiral boron reagent **28** is a more efficient promoter than **27**.^{54–56} The X-ray analysis of **28** demonstrates that it exists as a Brønsted acid-assisted chiral Lewis acid (BLA). **28** is prepared by the reaction of B(OMe)₃ (1 equiv) with (*R*)-binaphthol (2 equiv) in dichloromethane at reflux by removing methanol. Almost complete diastereoselectivities are obtained for a variety of imines under an optimum condition with a matching pair (Scheme 3.33).

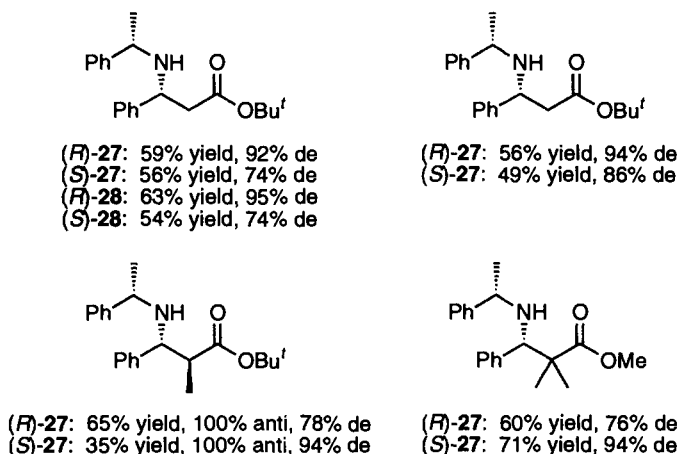


Scheme 3.33

3.3.3 Double asymmetric induction of aldol-type reaction of imines mediated by chiral boron reagents^{52–56}

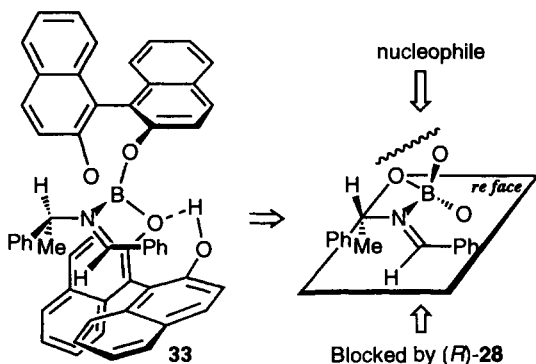
Largely stimulated by the synthesis of β -lactam antibiotics, there have been a great number of investigations into stereochemical aspects of imine condensations. We developed highly stereoselective aldol-type reaction of chiral imines derived from α -methylbenzylamine and aldehydes with ketene silyl acetals in the presence of chiral Lewis acid **27**^{54–56} or BLA **28**.^{52,53} Some examples are listed in Scheme 3.34. The matching and mismatching pairs in double stereo-differentiation depend on the structure of the ketene silyl acetals.

Examples



Scheme 3.34

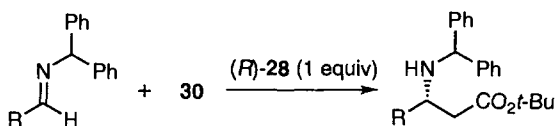
The absolute configuration of the adducts can be understood in terms of a rational model **33** involving an intramolecular hydrogen bonding interaction *via* Brønsted acid (Scheme 3.35).



Scheme 3.35

3.3.4 Enantioselective aldol-type reaction of imines mediated by BLA **28**^{52,53}

We developed the enantioselective synthesis of chiral β -amino acid esters from *achiral* imines and ketene silyl acetals using chiral BLA **28**. The enantioselectivity of the aldol-type reaction is dramatically increased by substituting sterically bulky *N*-substituents. The enantioselective aldol-type reaction of a variety of *N*-benzhydryl imines with **30** by (R)-**28** under optimum conditions is

Table 3.1 Enantioselective Aldo-Type Reaction of *N*-Benzhydrylimines

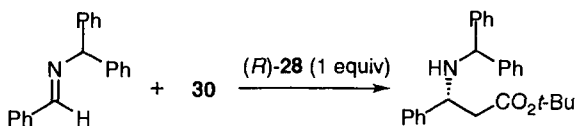
| Entry | R | Yield (%) | ee (%) |
|-------|---|-----------|-----------------|
| 1 | C ₆ H ₅ | 58 | 96 (<i>R</i>) |
| 2 | <i>p</i> -MeC ₆ H ₄ | 35 | 97 |
| 3 | <i>p</i> -ClC ₆ H ₄ | 45 | 98 |
| 4 | <i>p</i> -AcOC ₆ H ₄ | 52 | 98 |
| 5 | <i>p</i> -2,4-Cl ₂ C ₆ H ₃ | 49 | 95 |
| 6 | 2-Naphthyl | 43 | 96 |

summarized in Table 3.1. Excellent enantioselectivity (95% ee or better) has been achieved in the reactions of aromatic aldehyde-derived imines. The removal of *N*-benzhydryl protecting group from β -aryl β -amino acid esters is easily advanced by catalytic hydrogenation (10% Pd/C, H₂, MeOH).

Protocol 13.

Enantioselective aldol-type reaction of imines^{52,53} (Scheme 3.36)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 3.36**

Equipment

- Round-bottomed flask (10 mL)
- Separating funnel (50 mL)
- Magnetic stirrer bar
- Sintered glass filter funnel
- Magnetic stirrer
- Erlenmeyer flasks (50 mL)
- Column for flash chromatography
- Vacuum/argon source
- Pressure-equalized addition funnel
- Three-way stopcock

Materials

- (*R*)-(+)-1,1'-Bi-2-naphthol (FW 286.3), 200.4 mg, 0.70 mmol
- Trimethyl borate (FW 103.9), 0.1 M in CH₂Cl₂ 3.5 mL, 0.35 mmol
- 4Å molecular sieves
- *N*-Benzylidenebenzhydrylamine^{52,53} (FW 271.4), 95.0 mg

irritant

flammable liquid, moisture-sensitive
pellets
colourless crystal

3: Boron reagents

- **3 0** (FW 188.3) 131.8 mg, 0.7 mmol
- Dry dichloromethane^a
- Dry toluene^a

flammable
toxic, irritant
flammable liquid, toxic

1. Add (*R*)-binaphthol, trimethyl borate, and dichloromethane in a dry round-bottomed flask fitted with stirrer bar and a pressure-equalized addition funnel (containing a cotton plug and c. 4 g of 4Å molecular sieves and functioning as a Soxhlet extractor) surmounted by a reflux condenser.
2. Secure an argon atmosphere, and bring the solution to reflux (bath temperature 50–60°C).
3. After 2–3 h, cool the reaction mixture to 25°C, quickly remove the solution funnel and condenser and replace with a septum.
4. After adding 2 mL of dichloromethane, 5 mL of toluene, and *N*-benzylidenebenzhydramine to the white precipitate of (*R*)-**28** in dichloromethane at 0°C, stir the yellow suspension at 0°C for 10 min.
5. After cooling the suspension to –78°C, add **3 0** dropwise.
6. After stirring for 20 h, wash the solution with water and saturated NaHCO₃ and then dry over MgSO₄. Evaporation of the solvent and purification by column chromatography on silica gel gives *tert*-butyl 3-(benzhydramino)-3-phenylpropionate in 58% yield with 96% ee (*R*). The product is >98% pure by ¹H NMR, IR analysis, and may be characterized by elemental analysis. Determine the ee by HPLC analysis with a chiral Daicel OD-H column (hexane: *i*-PrOH = 100:1, 0.5 mL/min): *t*_R = 10.1 min for (*R*)-enantiomer, *t*_R = 12.1 min for (*S*)-enantiomer.

^a Distil dichloromethane and toluene from calcium hydride under argon.

References

1. Messey, A. G.; Park, A. J. *J. Organomet. Chem.* **1964**, 2, 245.
2. Messey, A. G.; Park, A. J. *J. Organomet. Chem.* **1966**, 5, 218.
3. Ishihara, K.; Hanaki, N.; Yamamoto, H. *Synlett* **1993**, 577.
4. Ishihara, K.; Funahashi, M.; Hanaki, N.; Miyata, M.; Yamamoto, H. *Synlett* **1994**, 963.
5. Ishihara, K.; Hanaki, N.; Funahashi, M.; Miyata, M.; Yamamoto, H. *Bull. Chem. Soc. Jpn.* **1995**, 68, 1721.
6. Ishihara, K.; Hanaki, N.; Yamamoto, H. *Synlett* **1996**, 721.
7. Parks, D. J.; Piers, W. E. *J. Am. Chem. Soc.* **1996**, 118, 9440.
8. Ishihara, K.; Kurihara, H.; Yamamoto, H. *Synlett* **1997**, 597.
9. Ishihara, K.; Ohara, S.; Yamamoto, H. *J. Org. Chem.* **1996**, 61, 4196.
10. Kaufmann, D.; Boese, R. *Angew. Chem., Int. Ed. Engl.* **1990**, 29, 545.
11. Furuta, K.; Miwa, Y.; Iwanaga, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1988**, 110, 6254.
12. Furuta, K.; Shimizu, S.; Miwa, Y.; Yamamoto, H. *J. Org. Chem.* **1989**, 54, 1481.

13. Gao, Q.; Yamamoto, H. *Org. Synth.* **1995**, 72, 86.
14. Furuta, K.; Kanematsu, A.; Yamamoto, H.; Takaoka, S. *Tetrahedron Lett.* **1989**, 30, 7231.
15. Ishihara, K.; Gao, Q.; Yamamoto, H. *J. Org. Chem.* **1993**, 58, 6917.
16. Ishihara, K.; Gao, Q.; Yamamoto, H. *J. Am. Chem. Soc.* **1993**, 115, 10412.
17. Sartor, D.; Saffrich, J.; Helmchen, G. *Synlett* **1990**, 197.
18. Takasu, M.; Yamamoto, H. *Synlett* **1990**, 194.
19. Sartor, D.; Saffrich, J.; Helmchen, G.; Richards, C. J.; Lambert, H. *Tetrahedron: Asymmetry* **1991**, 2, 639.
20. Corey, E. J.; Loh, T.-P. *J. Am. Chem. Soc.* **1991**, 113, 8966.
21. Corey, E. J.; Loh, T.-P.; Roper, T. D.; Azimioara, M. D.; Noe, M. C. *J. Am. Chem. Soc.* **1992**, 114, 8290.
22. Marshall, J. A.; Xie, S. *J. Org. Chem.* **1992**, 57, 2987.
23. Corey, E. J.; Guzman-perez, A. Loh, T.-P. *J. Am. Chem. Soc.* **1994**, 116, 3611.
24. Corey, E. J.; Loh, T.-P. *Tetrahedron Lett.* **1993**, 34, 3979.
25. Itsuno, S.; Kamahori, K.; Watanabe, K.; Koizumi, T.; Ito, K. *Tetrahedron: Asymmetry* **1994**, 5, 523.
26. Kamahori, K.; Tada, S.; Ito, K.; Itsuno, S. *Tetrahedron: Asymmetry* **1995**, 6, 2547.
27. Itsuno, S.; Watanabe, K.; Koizumi, T.; Ito, K. *React. Polym.* **1995**, 24, 219.
28. Kamahori, K.; Ito, K.; Itsuno, S. *J. Org. Chem.* **1996**, 61, 8321.
29. Hawkins, J. M.; Loren, S. *J. Am. Chem. Soc.* **1991**, 113, 7794.
30. Hawkins, J. M.; Loren, S.; Nambu, M. *J. Am. Chem. Soc.* **1994**, 116, 1657.
31. Kobayashi, S.; Murakami, M.; Harada, T.; Mukaiyama, T. *Chem. Lett.* **1991**, 1341.
32. Aggarwal, V. K.; Anderson, E.; Giles, R.; Zapparucha, A. *Tetrahedron: Asymmetry*, **1995**, 6, 1301.
33. Hayashi, Y.; Rohde, J.; Corey, E. J. *J. Am. Chem. Soc.* **1996**, 118, 5502.
34. Ishihara, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1994**, 116, 1561.
35. Ishihara, K.; Kurihara, H.; Yamamoto, H. *J. Am. Chem. Soc.* **1996**, 118, 3049.
36. Furuta, K.; Maruyama, T.; Yamamoto, H. *J. Am. Chem. Soc.* **1991**, 113, 1041.
37. Ishihara, K.; Maruyama, T.; Mouri, M.; Gao, Q.; Furuta, K.; Yamamoto, H. *Bull. Chem. Soc. Jpn* **1993**, 66, 3483.
38. Furuta, K.; Maruyama, T.; Yamamoto, H. *Synlett* **1991**, 439.
39. Furuta, K.; Mouri, M.; Yamamoto, H. *Synlett* **1991**, 561.
40. Ishihara, K.; Mouri, M.; Gao, Q.; Maruyama, T.; Furuta, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1993**, 115, 11490.
41. Marshall, J. A.; Tang, Y. *Synlett* **1992**, 653.
42. Kiyooka, S.; Kaneko, Y.; Komura, M.; Matsuo, H.; Nakano, M. *J. Org. Chem.* **1991**, 56, 2276.
43. Parmee, E. R.; Tempkin, O.; Masamune, S. *J. Am. Chem. Soc.* **1991**, 113, 9365.
44. Parmee, E. R.; Hong, Y. H.; Tempkin, O.; Masamune, S. *Tetrahedron Lett.* **1992**, 33, 1729.
45. Kiyooka, S.; Kaneko, T.Y.; Kume, K. *Tetrahedron Lett.* **1992**, 33, 4927.
46. Corey, E. J.; Cywin, C. L.; Roper, T. D. *Tetrahedron Lett.* **1992**, 33, 6907.
47. Kinugasa, M.; Harada, T.; Fujita, K.; Oku, A. *Synlett* **1996**, 43.
48. Kinugasa, M.; Harada, T.; Egusa, T.; Fujita, K.; Oku, A. *Bull. Chem. Soc. Jpn* **1996**, 69, 3639.
49. Kinugasa, M.; Harada, T.; Oku, A. *J. Org. Chem.* **1996**, 61, 6772.
50. Hattori, K.; Yamamoto, H. *J. Org. Chem.* **1992**, 57, 3264.

3: Boron reagents

51. Hattori, K.; Yamamoto, H. *Tetrahedron* **1993**, *49*, 1749.
52. Ishihara, K.; Miyata, M.; Hattori, K.; Yamamoto, H.; Tada, T. *J. Am. Chem. Soc.* **1994**, *116*, 10520.
53. Ishihara, K.; Kuroki, Y.; Yamamoto, H. *Synlett* **1995**, 41.
54. Hattori, K.; Miyata, M.; Yamamoto, H. *J. Am. Chem. Soc.* **1993**, *115*, 1151.
55. Hattori, K.; Yamamoto, H. *Synlett* **1993**, 239.
56. Hattori, K.; Yamamoto, H. *BioMed. Chem. Lett.* **1993**, *3*, 2337.

Magnesium(II) and other alkali and alkaline earth metals

AKIRA YANAGISAWA

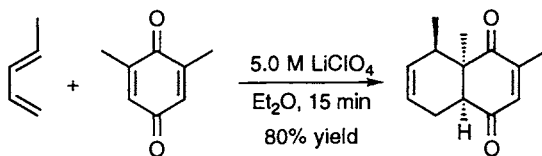
1. Introduction

Alkali and alkaline earth metal salts have mild Lewis acidities and have been utilized as a promoter or catalyst for organic reactions. Among the alkali metal salts, LiClO_4 ,¹⁻³ usually used as a concentrated solution in diethyl ether, is the most popular reagent to induce various transformations, including Diels–Alder reactions,^{4,5} 1,3-sigmatropic rearrangements,⁶ Mukaiyama aldol additions,^{7,8} and Michael additions.⁹ A catalytic amount of the lithium compound has also been found effective to promote these reactions.^{7,8,10} Additionally, LiBF_4 and LiNTf_2 have been used as lithium Lewis acid alternatives for Diels–Alder reactions.^{11,12} The alkaline earth metal salt MgBr_2 ¹³ is commonly used as a Lewis acid capable of forming strong bidentate chelates with α -alkoxy carbonyl compounds. It has thus been successfully applied to diastereoselective hetero-Diels–Alder reactions,^{14,15} Mukaiyama aldol additions,^{16,17} and allylation reactions.¹⁸⁻²¹ Chiral bis(oxazoline)-Mg(II) complexes have been developed and proven to be excellent chiral Lewis acid catalysts for enantioselective Diels–Alder reactions^{22,23} and 1,3-dipolar cycloaddition reactions.²⁴ The moderate Lewis acidity and reactivity of the chiral magnesium reagents are useful in obtaining high enantioselectivity and designing new catalysts for asymmetric synthesis.

2. Cycloaddition reactions

In 1986 Braun and Sauer found the enhanced endo selectivity for the Diels–Alder reaction of methyl acrylate and cyclopentadiene in concentrated solutions of lithium perchlorate in diethyl ether, THF, and DME.⁴ Four years later Grieco *et al.* described such a solvent system, a 5.0 M solution of LiClO_4 in diethyl ether, which had a greater accelerating effect on the [4+2]-cycloaddition.⁵ For example, the reaction of *trans*-piperylene with 2,6-dimethylbenzoquinone in a 5.0 M ether solution of LiClO_4 at ambient temperature and pressure is complete within 15 min, affording an 80% yield of cycloadduct

(Scheme 4.1), whereas the same reaction in benzene requires a longer reaction time and higher reaction temperature for completion of the reaction. The observed rate acceleration, which has been believed to come from a high internal solvent pressure, is due to the Lewis acidity of the lithium salt.²⁵ In 1993 Reetz *et al.* found that a catalytic amount of LiClO_4 (3–25 mol%) suspended in dichloromethane is effective in promoting Diels–Alder reactions¹⁰ in addition to Mukaiyama aldol reactions,⁸ Michael additions,⁸ and 1,3-Claisen rearrangements.¹⁰ An example is provided by the reaction of methyl acrylate with cyclopentadiene in the presence of 25 mol% of LiClO_4 (Scheme 4.2 in Protocol 1).¹⁰ The conversion is 87% at -15°C after a reaction time of 24 h and the *endo:exo* ratio is shown to be 7.3:1.

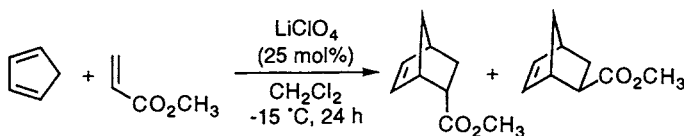


Scheme 4.1

Protocol 1.

Diels–Alder reaction of methyl acrylate with cyclopentadiene catalysed by LiClO_4 ¹⁰ (Scheme 4.2)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 4.2

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice–acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Materials

- Cyclopentadiene, 330 mg, 5 mmol
- Methyl acrylate, 86 mg, 1 mmol
- Lithium perchlorate, 25 mg, 0.25 mmol
- Dichloromethane, 16 mL
- Water
- Magnesium sulfate

flammable, toxic
flammable, lachrymator
oxidizer, irritant
toxic, irritant
hygroscopic

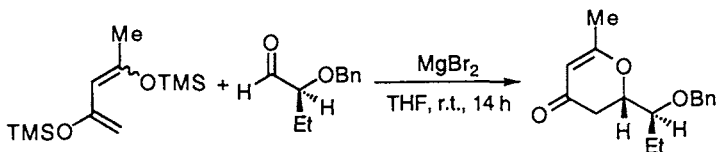
1. Ensure that the Schlenk flask and syringes have been dried and flushed with argon before use.
2. Dry dichloromethane by distillation over P_2O_5 and then CaH_2 prior to use.
3. Prepare a solution of methyl acrylate (86 mg, 1 mmol) and cyclopentadiene (330 mg, 5 mmol) in dry CH_2Cl_2 (1 mL) under argon atmosphere.
4. After cooling the solution to $-15^\circ C$, add $LiClO_4$ (25 mg, 0.25 mmol).
5. After stirring for 24 h, dilute the reaction mixture with CH_2Cl_2 (15 mL).
6. Wash the dichloromethane solution with water twice (15 mL each). Dry the organic layer and then remove the organic solvent using a rotary evaporator to obtain the crude product (87% conversion, an *endo:exo* ratio of 7.3:1).
7. Determine the conversion based on methyl acrylate and the *endo:exo* ratio of the crude product by GC analysis.

Magnesium halides have been utilized as Lewis acids for [4+2]-cycloaddition reactions and shown to increase the regioselectivity, having an opposite effect to that of $BF_3 \cdot OEt_2$.^{14,15,26} For example, the $MgBr_2$ -promoted hetero-Diels–Alder reaction of a siloxy diene with α -alkoxy aldehydes gives a single diastereomer (Scheme 4.3 in Protocol 2),¹⁵ whereas under Lewis acid catalysis by either $BF_3 \cdot OEt_2$ or $ZnCl_2$, another diastereomer is selectively obtained with a 1.4:1 ratio. The stereochemical outcome of the $MgBr_2$ -promoted reaction reveals that the cycloaddition occurs from the less hindered face of a *syn* conformer of the aldehyde in which a chelation of the magnesium exists between the two oxygens.

Protocol 2.

Hetero-Diels–Alder reaction of 2,4-Bis(trimethylsiloxy)-1,3-pentadiene with 2-(benzyloxy)butyraldehyde promoted by magnesium bromide. synthesis of (2*S,1'*S**)-2-[(1'-benzyloxy)propyl-6-methyl-2,3-dihydro-4*H*-pyran-4-one]¹⁵ (Scheme 4.3)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 4.3

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with ice water
- Round-bottomed flask (200 mL)
- Vacuum/inert gas source (argon)

Protocol 2. Continued**Materials**

| | |
|---|--------------------------------------|
| • 2,4-Bis(trimethylsiloxy)-1,3-pentadiene, 9.10 g, 42.1 mmol | flammable, moisture-sensitive |
| • 2-(Benzyloxy)butyraldehyde, 3.00 g, 16.9 mmol | irritant |
| • Magnesium bromide (2.95 M solution in 10% benzene/ether), ^a 5.7 mL, 16.9 mmol | irritant, hygroscopic |
| • Dry tetrahydrofuran (THF), 80 mL | flammable, irritant |
| • Sodium bicarbonate | moisture-sensitive |
| • Ethyl ether | flammable, irritant |
| • Magnesium sulfate | hygroscopic |
| • Dichloromethane, 100 mL | toxic, irritant |
| • Trifluoroacetic acid, 4 mL | corrosive, toxic |
| • Sodium chloride | hygroscopic, irritant |

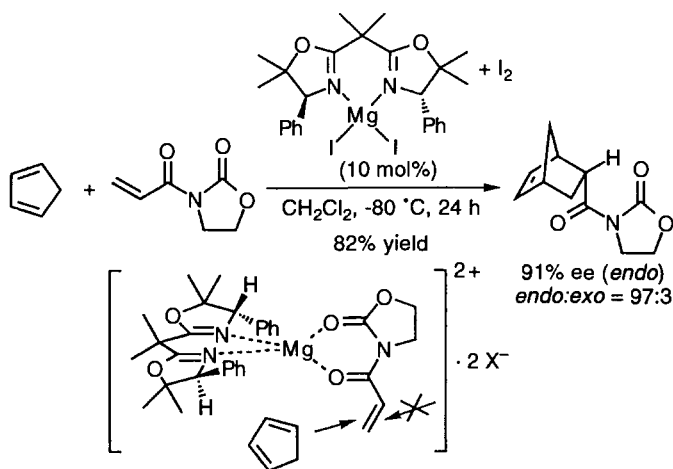
1. Ensure that all glass equipment has been dried in an oven before use.
2. Place 2-(benzyloxy)butyraldehyde (3.00 g, 16.9 mmol) in a 200 mL flask with a magnetic stirring bar under argon atmosphere and dissolve the aldehyde with dry THF (80 mL).
3. Cool the solution to 0°C and add a solution of magnesium bromide (2.95 M, 5.7 mL, 16.9 mmol) in 10% benzene/ether over 2 min.
4. After stirring for 10 min, add 2,4-bis(trimethylsiloxy)-1,3-pentadiene (9.10 g, 42.1 mmol) and warm the solution slowly to room temperature.
5. After stirring for 14 h, pour the solution into a saturated aqueous sodium bicarbonate solution and extract with ether. Dry the organic layer over anhydrous MgSO₄ and concentrate *in vacuo*.
6. Dissolve the resultant oil in 100 mL of CH₂Cl₂ and add trifluoroacetic acid (4 mL). After stirring for 3 h, wash the solution with water, saturated aqueous sodium bicarbonate solution, and brine, then dry the organic layer over anhydrous MgSO₄ and concentrate *in vacuo*.
7. Purify the crude product by flash column chromatography on silica gel (40% ethyl acetate/hexane) to give 3.41 g (78%) of the title compound as a clear colourless oil and characterize by IR, ¹H NMR, ¹³C NMR, and elemental analysis.
8. Examine the crude and purified products by GLC (column: 4 ft, 3% OV-17, 210°C, 30 mL/min flow rate, *t_R* = 10.1 min) and 500 MHz ¹H NMR experiments to confirm there is no detectable amount of the diastereomer.

^a Magnesium bromide can be prepared from 1,2-dibromoethane and magnesium turnings in ether. Add benzene (c. 10%) to homogenize the two-phase mixture. Calculate the molarity by weighing the residual magnesium. The resultant solution can be stored for several months at room temperature.

In 1992 the first example of a chiral magnesium reagent as a catalyst for an asymmetric reaction was reported by Corey and Ishihara.²² The chiral Mg catalyst is prepared by treatment of a chiral bis(oxazoline) ligand with MgI₂ and I₂ (co-catalyst) in CH₂Cl₂. When 3-acryloyloxazolidine-2-one is reacted with cyclopentadiene in the presence of the catalyst (10 mol%) at -80°C, the Diels-Alder product was obtained in 82% yield with 91% ee and an *endexo*

4: Magnesium(II) and other alkali and alkaline earth metals

ratio of 97/3 (Scheme 4.4).²² A 1:1:1 complex of the chiral ligand, Mg^{2+} , and the dienophile is believed to be the reactive species in which the *s-cis* form of oxazolidinone co-ordinates to the metal. The diene approaches from the front side of the complex avoiding repulsion by the phenyl group (Scheme 4.4).



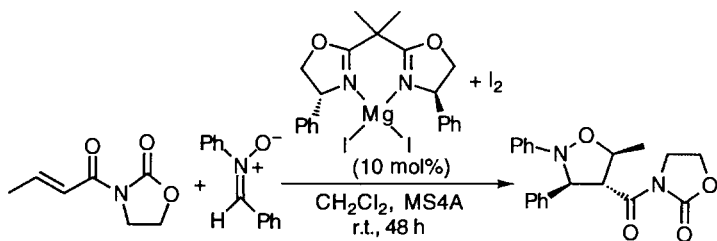
Scheme 4.4

This chiral bis(oxazoline)-magnesium catalyst has been successfully applied to the enantioselective 1,3-dipolar cycloaddition reaction of alkenes with nitrones to lead high *endo*-selectivity with up to 82% ee (Scheme 4.5 in Protocol 3).²⁴

Protocol 3.

Asymmetric 1,3-dipolar cycloaddition reaction of 3-((*E*)-2'-butenyl)-1,3-oxazolidin-2-one with benzyldenephénylamine *N*-oxide catalysed by chiral Mg(II)-bisoxazoline complex²⁴ (Scheme 4.5)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 4.5

Protocol 3. Continued**Equipment**

- Magnetic stirrer
- Syringes
- Schlenk flask (10 mL, 5 mL)
- Vacuum/inert gas source (argon)

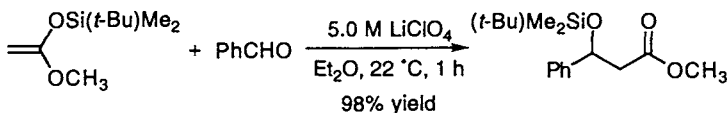
Materials

- 3-((*E*)-2'-butenoyl)-1,3-oxazolidin-2-one, 78 mg, 0.5 mmol **flammable, toxic**
- Benzyldenenephénylamine *N*-oxide, 123 mg, 0.625 mmol
- (*R*)-2,2'-isopropylidenebis(4-phenyl-2-oxazoline), 410 mg, 1.25 mmol
- Magnesium, 48 mg, 2.0 mmol **moisture-sensitive**
- Iodine, 506 mg, 2.0 mmol **toxic, corrosive**
- Ethyl ether, 5 mL **flammable, irritant**
- Dichloromethane, 15 mL **toxic, irritant**
- 4Å powdered molecular sieves, 50–100 mg
- MeOH

1. Ensure that all glass equipment has been dried in an oven before use.
2. Activate the 4Å powdered molecular sieves by heating to 250°C for 3 h in high vacuum prior to use.
3. Place magnesium (48 mg, 2.0 mmol) and iodine (253 mg, 1.0 mmol) in a 10 mL flask with a magnetic stirring bar under N₂. Add diethyl ether (5 mL) and stir the mixture at room temperature until the iodine colour disappears (2–3 h). Filter off the unreacted Mg and remove the solvent of the filtrate under reduced pressure at room temperature.
4. Dissolve the white MgI₂ in CH₂Cl₂ and add (*R*)-2,2'-isopropylidenebis(4-phenyl-2-oxazoline) (410 mg, 1.25 mmol) dissolved in CH₂Cl₂ (5 mL) to the resulting solution.
5. After stirring the milky suspension for 2 h, add I₂ (253 mg, 1.0 mmol) and stir the deep-red suspension for 2 h before use.
6. Add 3-((*E*)-2'-butenoyl)-1,3-oxazolidin-2-one (78 mg, 0.5 mmol) and benzyldenenephénylamine *N*-oxide (123 mg, 0.625 mmol) to a suspension of 50–100 mg of 4Å powdered molecular sieves in CH₂Cl₂ (2 mL) in a 5 mL flask.
7. After stirring the mixture for 15 min, add the catalyst (0.05 mmol) with a glass pipette.
8. After stirring for 48 h, treat the reaction mixture with 5% MeOH in CH₂Cl₂ (1 mL) and filter the mixture through a 20 mm layer of silica gel.
9. Wash the silica gel layer with another 2 mL of 5% MeOH in CH₂Cl₂ and remove the organic solvent using a rotary evaporator.
10. Purify the crude product by preparative TLC (silica gel, ether–petroleum ether (3:1)) to give a pure product in 81% yield (84% *endo*, 75% ee).
11. Determine the enantiomeric excess by HPLC analysis (Daicel Chiralcel OD, hexane/isopropanol = 9:1, flow rate = 1.0 mL/min), *t_R* = 42 min (minor), *t_R* = 58 min (major).

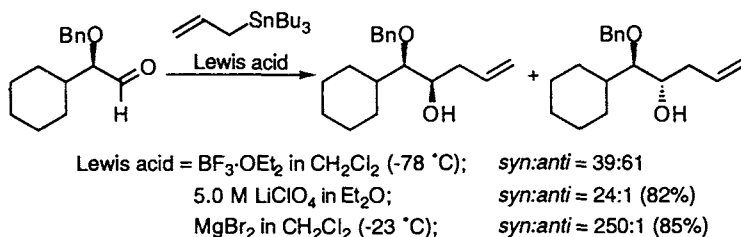
3. Aldol additions, allylations, and other reactions

Lithium and magnesium Lewis acid reagents are also effective promoters for nucleophilic carbon-carbon bond forming reactions. The Mukaiyama aldol reaction of ketene silyl acetals with aldehydes occurs at room temperature under the influence of a stoichiometric or catalytic amount of LiClO_4 .^{7,8} For example, the silyl ether derived from methyl acetate reacts with benzaldehyde in a 5.0 M ether solution of LiClO_4 to provide the silylated aldol adduct in 98% yield (Scheme 4.6).⁷ The use of a catalytic amount (3 mol%) of LiClO_4 suspended in CH_2Cl_2 results in complete conversion within 15 min.⁸



Scheme 4.6

The addition of allylstannanes to α -alkoxy aldehydes has been shown to proceed with high diastereoselectivity under chelation control in the presence of LiClO_4 ²⁷ or MgBr_2 .^{18–21} When a mixture of α -benzyloxy aldehyde and allyltributylstannane is treated with MgBr_2 in CH_2Cl_2 , the *syn*-product is obtained with >250:1 stereoselectivity.¹⁸ In marked contrast, non-chelation-controlled stereochemistry is observed with $\text{BF}_3 \cdot \text{OEt}_2$ (Scheme 4.7).¹⁸ The high *syn*-selectivity provided by MgBr_2 is due to the formation of a discrete bidentate chelate with the α -alkoxy aldehyde.



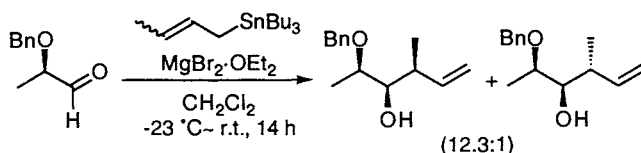
Scheme 4.7

The high levels of diastereofacial selectivity observed for the allylations are preserved for the case of crotylation where three contiguous chiral centres are defined during the reaction.^{19,20} With α -benzyloxy aldehyde, *syn,syn*-selectivity is observed regardless of the geometry of the crotyltin. However, higher diastereoselectivities are given with the *E*-stannane, for instance, a 12.3:1 ratio of *syn,syn/syn,anti* products is obtained with a 90:10 mixture of the *E*- and *Z*-stannanes (Scheme 4.8 in Protocol 4).²¹

Protocol 4.

Addition of crotyltributylstannane to (*R*)-2-(benzyloxy)propionaldehyde promoted by magnesium bromide etherate. Preparation of (*2R,3R,4S*)-2-benzyloxy-4-methyl-5-hexen-3-ol²¹ (Scheme 4.8)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 4.8

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice/acetone
- Round-bottomed flask (100 mL)
- Vacuum/inert gas source (argon)

Materials

- | | |
|---|---|
| <ul style="list-style-type: none"> • (<i>R</i>)-2-(Benzyloxy)propionaldehyde, 450 mg, 2.74 mmol • Crotyltributylstannane (<i>E:Z</i> = 90:10),^a 1.42 g, 4.11 mmol • Magnesium bromide etherate, 1.14 g, 5.48 mmol • Dry dichloromethane, 30 mL • Sodium bicarbonate • Dichloromethane for extraction (80 mL) • Sodium sulfate | <p>irritant</p> <p>irritant</p> <p>flammable, moisture-sensitive</p> <p>toxic, irritant</p> <p>moisture-sensitive</p> <p>toxic, irritant</p> <p>irritant, hygroscopic</p> |
|---|---|

1. Ensure that all glass equipment has been dried in an oven before use.
2. Place (*R*)-2-(benzyloxy)propionaldehyde (450 mg, 2.74 mmol) in an oven-dried 100 mL round bottom flask with a magnetic stirring bar under argon atmosphere and dilute the aldehyde with dry dichloromethane (30 mL).
3. After cooling the solution to -23°C for 10 min, add magnesium bromide etherate (1.14 g, 5.48 mmol) all at once as a white powder. The mixture will become cloudy, and then over the next 10 min the solid material will partially dissolve and the solution will become slightly yellow.
4. After stirring for 15 min, add crotyltributylstannane (1.42 g, 4.11 mmol) dropwise via a syringe down the side of the flask to allow for ample cooling. Stir the mixture for an additional 2 h at -23°C before the bath is allowed to expire.
5. After stirring for 12 h at room temperature, add saturated aqueous sodium bicarbonate solution (30 mL) and continue stirring for 25 min. Extract the mixture with CH_2Cl_2 (4×20 mL) and dry the organic layer over anhydrous Na_2SO_4 . Filter the organic layer through a plug of Celite[®] (1 cm) and silica gel (2 cm).

- Remove the solvent *in vacuo* and purify the crude product via radial plate liquid chromatography. Load the material onto a 4 mm plate with hexane (3 mL) and elute with 150 mL of hexane, 100 mL of 5% EtOAc/hexane, 100 mL of 10% EtOAc/hexane, and 150 mL of 15% EtOAc/hexane to give the title compound (509 mg, 84%, (2*R*,3*R*,4*S*)-isomer:(2*R*,3*R*,4*R*)-isomer = 12.3:1) as a clear colourless liquid.
- Determine the diastereomeric ratio by GLC analysis (GC DX-4 15 m column 120–180 °C at 3.5 °C/min, major, 13.72 min, minor, 14.00 min).

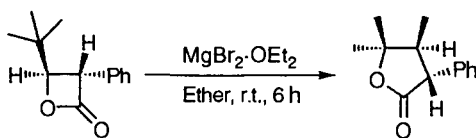
^a Crotyltributylstannane can be obtained by tin anion displacement of the corresponding crotyl chloride.

The oxophilic nature of MgBr₂ is valid for mediating numerous rearrangements of oxygen-containing compounds. A typical example of the stereospecific conversion of a β -lactone to a butyrolactone including methyl group migration is shown in Scheme 4.9 (Protocol 5).²⁸ Use of other conventional Lewis acids than MgBr₂•OEt₂ results in recovery of the unreacted β -lactone or the occurrence of undesired side reactions. A stepwise mechanism involving a rate-determining ionization step is proposed for the rearrangement rather than a concerted mechanism.²⁹

Protocol 5.

Rearrangement of a β -lactone to a butyrolactone promoted by magnesium bromide etherate. Preparation of *trans*-4,5-dihydro-3-phenyl-4,5,5-trimethyl-2(3*H*)-furanone²⁹ (Scheme 4.9)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 4.9

Equipment

- Magnetic stirrer
- Syringes
- Three-necked flask (25 mL)
- Vacuum/inert gas source (nitrogen)

Materials

- *cis*-4-*tert*-Butyl-3-phenyloxetan-2-one,^a 2.04 g, 10 mmol
- Magnesium bromide etherate, 2.58 g, 10 mmol
- Dry ether,^b 10 mL
- Water, 10 mL
- Magnesium sulfate
- Hexane for recrystallization

irritant
flammable, moisture-sensitive
flammable, irritant

hygroscopic
flammable, irritant

Protocol 5. Continued

1. Ensure that all glass equipment has been dried in an oven at 120°C for a minimum of 4 h and then assembled under a nitrogen stream before use.
2. Equip an oven-dried 25-mL three-necked flask with a nitrogen inlet and stirring bar and charge the flask with 10 mL of a 1.0 M solution of *cis*-4-*tert*-butyl-3-phenyloxetan-2-one (2.04 g, 10 mmol) in anhydrous ether.
3. Begin stirring and add magnesium bromide etherate (2.58 g, 10 mmol) in a single portion. Stir the light-yellow mixture under nitrogen for 6 h at room temperature, then terminate the reaction by the gradual addition of 10 mL of water.
4. Separate the layers and dry the ether layer over anhydrous MgSO₄. After filtering the organic layer, remove the solvent under reduced pressure. Purify the crude product by recrystallization from hexane to afford the title compound (1.75 g, 86%) which displays the appropriate ¹H NMR (in CDCl₃) and IR (KBr).

^a *cis*-4-*tert*-Butyl-3-phenyloxetan-2-one can be conveniently prepared from phenylacetic acid by a two-step sequence: (1) deprotonation of the acid by 2 equiv of lithium di-isopropylamide in THF followed by treatment with 2,2-dimethylpropanal to afford the resulting β-hydroxy acid (74%), (2) cyclization of the acid with benzenesulfonyl chloride in pyridine at 0°C (94%).

^b Dry ether is distilled from sodium benzophenone ketyl prior to use.

References

1. Grieco, P. A. *Aldrichimica Acta* **1991**, 24, 59.
2. Waldmann, H. *Angew. Chem. Int. Ed. Engl.* **1991**, 30, 1306.
3. Charette, A. B. In *Encyclopedia of reagents for organic synthesis* (ed. L. A. Paquette), John Wiley, Chichester, **1995**, p 3155.
4. Braun, R.; Sauer, J. *Chem. Ber.* **1986**, 119, 1269.
5. Grieco, P. A.; Nunes, J. J.; Gaul, M. D. *J. Am. Chem. Soc.* **1990**, 112, 4595.
6. Grieco, P. A.; Clark, J. D.; Jagoe, C. T. *J. Am. Chem. Soc.* **1991**, 113, 5488.
7. Reetz, M. T.; Raguse, B.; Marth, C. F.; Hügel, H. M.; Bach, T.; Fox, D. N. A. *Tetrahedron* **1992**, 48, 5731.
8. Reetz, M. T.; Fox, D. N. A. *Tetrahedron Lett.* **1993**, 34, 1119.
9. Grieco, P. A.; Cooke, R. J.; Henry, K. J.; VanderRoest, J. M. *Tetrahedron Lett.* **1991**, 32, 4665.
10. Reetz, M. T.; Gansäuer, A. *Tetrahedron* **1993**, 49, 6025.
11. Smith, D. A.; Houk, K. N. *Tetrahedron Lett.* **1991**, 32, 1549.
12. Handy, S. T.; Grieco, P. A.; Mineur, C.; Ghosez, L. *Synlett* **1995**, 565.
13. Black, T. H. In *Encyclopedia of reagents for organic synthesis* (ed. L. A. Paquette), John Wiley, Chichester, **1995**, p 3197.
14. Danishefsky, S. J.; Pearson, W. H.; Harvey, D. F. *J. Am. Chem. Soc.* **1984**, 106, 2455, 2456.
15. Danishefsky, S. J.; Pearson, W. H.; Harvey, D. F.; Maring, C. J.; Springer, J. P. *J. Am. Chem. Soc.* **1985**, 107, 1256.

4: Magnesium(II) and other alkali and alkaline earth metals

16. Bernardi, A.; Cardani, S.; Colombo, L.; Poli, G.; Schimperna, G.; Scolastico, C. *J. Org. Chem.* **1987**, 52, 888.
17. Annunziata, R.; Cinquini, M.; Cozzi, F.; Cozzi, P. G.; Consolandi, E. *J. Org. Chem.* **1992**, 57, 456.
18. Keck, G. E.; Boden, E. P. *Tetrahedron Lett.* **1984**, 25, 265.
19. Keck, G. E.; Boden, E. P. *Tetrahedron Lett.* **1984**, 25, 1879.
20. Keck, G. E.; Abbott, D. E. *Tetrahedron Lett.* **1984**, 25, 265, 1883.
21. Keck, G. E.; Savin, K. A.; Cressman, E. N. K.; Abbott, D. E. *J. Org. Chem.* **1994**, 59, 7889.
22. Corey, E. J.; Ishihara, K. *Tetrahedron Lett.* **1992**, 33, 6807.
23. Desimoni, G.; Faita, G.; Righetti, P. P. *Tetrahedron Lett.* **1996**, 37, 3027.
24. Gothelf, K. V.; Hazell, R. G.; Jørgensen, K. A. *J. Org. Chem.* **1996**, 61, 346.
25. Forman, M. A.; Dailey, W. P. *J. Am. Chem. Soc.* **1991**, 113, 2761.
26. Kelly, T. R.; Montury, M. *Tetrahedron Lett.* **1978**, 4311.
27. Henry, K. J., Jr.; Grieco, P. A.; Jagoe, C. T. *Tetrahedron Lett.* **1992**, 33, 1817.
28. Black, T. H.; Hall, J. A.; Sheu, R. G. *J. Org. Chem.* **1988**, 53, 2371.
29. Black, T. H.; Eisenbeis, S. A.; McDermott, T. S.; Maluleka, S. L. *Tetrahedron* **1990**, 46, 2307.

Zinc(II) Reagents

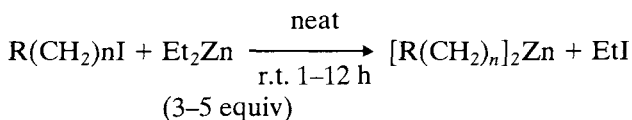
NOBUKI OGUNI

1. Introduction

There are two kinds of organozinc compounds: Grignard-type reagents and dialkylzincs. Organozinc reagents have been used for the carbon-carbon bond-forming reactions, for example, Reformatsky and Simmons-Smith reactions. Nevertheless dialkylzinc compounds have been little used for organic reactions except organometallic component of Ziegler catalyst. After our first report of the discovery in 1883 on the enantioselective addition of diethylzinc to benzaldehyde catalyzed by chiral β -aminoalcohols, this field of asymmetric carbon-carbon bond forming catalytic reaction was developed largely by many researchers. Also the asymmetric amplification, that is extreme amplification in enantioselectivity was found in this reaction and gave the extensive effect for all enantioselective catalytic reactions.¹⁻⁴

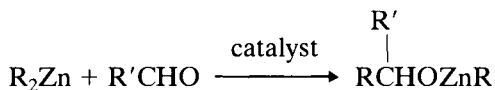
2. Scope and limitations

Dialkylzincs can be prepared easily by transmetallation from alkylaluminiums and alkylboranes. Particularly, diethylzinc and dimethylzinc among many dialkylzincs are commercially available in most countries. Therefore the synthetic method of non-commercially available organozinc compounds using diethylzinc will be convenient and advantageous in comparison with other methods.^{3,4}



Dialkylzincs have been known to be unreactive to aldehydes, esters, imines, and so on, in polar solvents at room temperature. This characteristic of

organozincs is quite different from alkyl-lithium and dialkylmagnesium which react easily on mixing with aldehydes, ketones, and other functionalized compounds even at low temperature. Dialkylzinc only reacts with aldehydes to give addition products in the presence of basic catalysts such as trialkyl amines. The reaction products are monoalkylzinc alkoxides which easily associate to afford cubic tetramers co-ordinated with oxygen atoms of alkoxides. The alkyl group of alkylzinc alkoxide does not react with aldehydes.

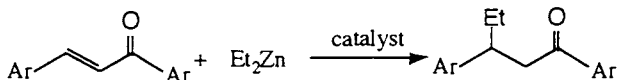


Dialkylzincs also react with β -aminoalcohols or β -thioalcohols to give 2-valence 3-co-ordinated alkylzinc compounds which also associate to form dimers. When aldehydes and excess dialkylzincs are present in the media, dimer easily dissociate to catalyse the addition reaction of dialkylzincs to aldehydes. There have been reports of the use of chiral β -aminoalcohols⁵⁻⁵⁰ containing β -diamines, β -aminothiols,⁶⁸⁻⁷² hydroxy-pyridines⁵¹⁻⁵⁷ and amino-alcohol derivatives of ferrocene⁵⁸⁻⁶³ and benzene-chrom complexes⁶⁴⁻⁷² in the reaction media of dialkylzincs with aldehydes to afford optically active products. Generally the chirality of the carbon bonded with a hydroxy group affects predominantly the chirality of the product rather than one with an amino group. In general aromatic and α,β -unsaturated aldehydes give high enantioselective products in reactions using β -difunctionalized chiral auxiliaries. In enantioselective addition of dialkylzincs to aldehydes, asymmetric amplification was found, even if a small amount of β -aminoalcohols bearing low optical purity was used as a catalyst, very high enantiomeric products were obtained. Investigations of the mechanism of this curious reaction estimated that the monomeric zincaminoalkoxide is the real active catalyst, and its racemic dimer associate cannot dissociate in reaction media.⁷³⁻⁷⁶ When the chemical constitution and absolute configuration of the product are exactly the same with chiral auxiliary used, the product itself can also catalyse the reaction to produce the same compound. This reaction was realized in isopropylation of pyridinecarbaldehydes with di-isopropylzinc, by a reaction called autocatalysis.⁷⁷⁻⁸²

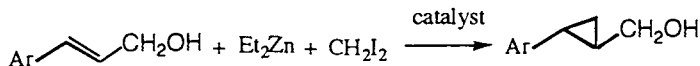
Chiral Lewis acid-catalysed alkylation of aldehydes with dialkylzincs gave highly enantioselective products. Two outstanding catalysts are the titanium-chiral TADDOL and titanium-chiral disulfonamido systems.⁸³⁻¹⁰⁵ Both catalysts give the very high enantiomeric products from all kinds of aldehydes, both aliphatic and alicyclic. However, in the former catalyst system an equimolar amount of titanium alkoxides to aldehydes is needed to obtain addition products in high yield.

Highly enantioselective 1,4-addition of diethylzinc to α,β -unsaturated ketones was attained using nickel or copper complexes of β -aminoalcohols.¹⁰⁶⁻¹¹⁰ Also the catalytic enantioselective cyclopropanation of allylic alcohols with

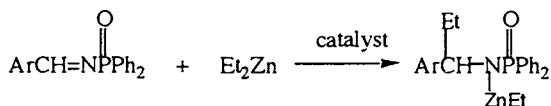
5: Zinc(II) Reagents



diethylzinc- CH_2I_2 was found using chiral titanium-disulfonamido complexes as catalyst.^{111–124}



Dialkylzinc can add to imino groups by catalysis of chiral titanium complexes to give chiral amine derivatives, followed by hydrolysis to afford optical active primary amines.^{125–129}

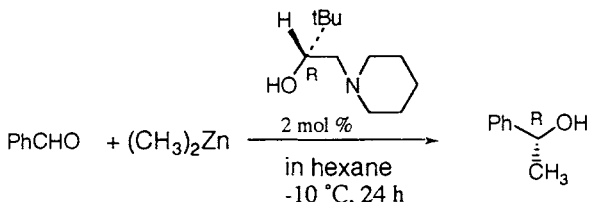


Cautions: Dialkylzincs are highly flammable liquids, which induce spontaneous ignition in air, and react violently with water and alcohols. Therefore, dialkylzincs should be kept under an inert atmosphere.

Protocol 1.

Chiral β -aminoalcohol-mediated asymmetric addition of diethylzinc to benzaldehyde⁸ (Scheme 5.1)

Alkylation, particularly ethylation of aromatic compounds has been investigated with many kinds of β -aminoalcohols and highly enantioselective reactions also reported.



Scheme 5.1

Equipment

- Two-necked, round-bottomed flask (500 mL) fitted with rubber septum and a Teflon-coated magnetic stirring bar
- Three-way stopcock fitted to the top of flask and connected to a vacuum/argon (or dry nitrogen) source
- Magnetic stirrer
- All-glass syringe with a needle-lock Luer and medium-gauge needle
- Water-jacketed short-path distillation apparatus
- Separating funnel (500 mL)
- Rotary evaporator

Protocol 1. Continued**Materials**

- Dry toluene
- Benzaldehyde
- Diethylzinc
- (*R*)-1-*t*-Butyl-2-piperidinoethanol
- Technical ether for extraction

explosive on contact to air**flammable**

1. Flame dry the reaction vessel and place with a stirrer bar *in vacuo*. Fill the flask with argon and maintain a slightly positive argon pressure throughout this reaction.
2. Assemble the syringe and needle while hot and flush the syringe with argon.
3. Support the flask using clamp and a stand with a heavy base.
4. Charge the flask with toluene (200 mL) using a syringe by piercing the septum on the reaction flask.
5. Keep the flask at -40°C with a dry ice/acetone bath.
6. Add benzaldehyde (10 g, 94.2 mmol) and (*R*)-1-*t*-butyl-2-piperidinoethanol (0.37 g, 2.0 mmol) using syringes under stirring.
7. Add diethylzinc (10.5 mL, 102 mmol) dropwise from the syringe through the septum on the reaction flask to the mixture in the flask with vigorous stirring.
8. Allow the temperature of the flask to rise from -40°C to -10°C with an ice/NaCl bath, and keep the mixture at -10°C for 24 h.
9. Quench the mixture by dropping 2M HCl (100 mL) with stirring at 0°C .
10. Transfer the mixture to a separating funnel and separate the two layers. Extract the water layer with ether (2×50 mL).
11. Transfer the organic layer to a 500 mL flask. Dry the organic layer over Na_2SO_4 , and filter through a filter paper. Concentrate the filtrate under reduced pressure using a rotary evaporator ($30^{\circ}\text{C}/30$ mm Hg).
12. Transfer the oily residue to a water-jacketed, short path distillation apparatus equipped with a thermometer. Distil the crude product under reduced pressure to obtain 1-phenyl-1-propanol(b.p. $103\text{--}104^{\circ}\text{C}$ 14 mm Hg; 12.6 g, 98% yield) as a colourless oil.
13. Determine the absolute configuration and enantiomeric excess(*ee*) of the product. The absolute configuration can be determined by the sign of $[\alpha]_{\text{D}}$ $[\alpha]_{\text{D}} -45.4$.(*c* 2.0, $\text{C}_2\text{H}_5\text{OH}$) indicated a (*S*)-configuration. *Ee* ($>98\%$) was determined by HPLC analysis (column, Daicel CHIRALCEL OB; 100/0.2 hexane/2-propanol; t_{R} of (*S*)-isomer, 12 min; t_{R} of (*R*)-isomer, 13 min.

(*R*)-1-Dialkylamino-3-3-dimethylbutan-2-ol was prepared by the reaction of (*R*)-*t*-butyl-ethylene oxide^a with bromomagnesium dialkylamide as follows.

To a solution of ethylmagnesium bromide (5.0 mmol) in tetrahydrofuran (THF)(10 mL) was added a solution of a secondary amine (5.0 mmol) in THF(5

mL). After the mixture has been stirred at 35°C for 1 h, (*R*)-*t*-butylethylene oxide (500 mg, 5 mmol) was added to the solution, which was then stirred for 6 h at room temperature before being poured into saturated aq.NH₄Cl. The aqueous solution was acidified with 1 mol dm⁻³ HCl (20 ml), extracted with ethyl acetate (30 mL × 2), and then made alkaline by 10% aq. NaOH and extracted with ethyl acetate (20 mL × 2). The combined organic layer was washed with brine (20 mL × 2), dried over Na₂SO₄, and distilled or recrystallized to give the corresponding (*R*)-1-dialkylamino-3,3-dimethylbutan-2-ol in good yield. The enantiomeric excess of the β-aminoalcohol thus obtained was determined as over 99% by HPLC analysis of the corresponding 3,5-dinitrophenyl carbamate, using a chiral stationary phase(column, Sumitomo Chemical Co. Sumipax OA 4000).

(*R*)-3,3-Dimethyl-1-piperidinobutan-2-ol (**1**) was obtained from the reaction of piperidine (425 mg, 5 mmol) and (*R*)-*t*-butylethylene oxide (500 mg, 5 mmol) in yield of 743 mg (80%), b.p. 59–61°C/1 mmHg). [α]_D²⁵ –72.4 (C 1.8, CHCl₃).

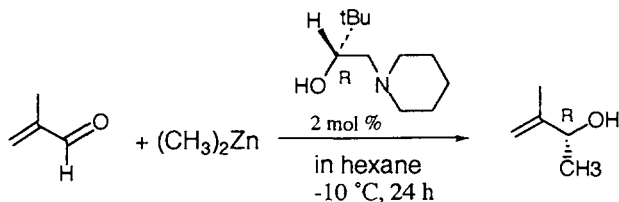
(*R*)-1-(3-Azabicyclo[3.2.2]nonan-3-yl)-3-dimethylbutan-2-ol was obtained from the reaction of 3-azabicyclo[3.3.3]nonane (630 mg, 5 mmol) and (*R*)-*t*-butylethylene oxide (500 mg, 5 mmol) in yield of 1.12 g (88%), m.p.105–108°C). [α]_D²² –64.2° (C 1.0, CHCl₃).

^a Leven, P. A.; Walti, A. Organic Synthesis **1943**, Coll. Vol. 2, 5–9. Hurst, S. J.; Bruce, J. M. *J. Chem. Soc.* **1963**, 147.

Protocol 2.

Chiral β-aminoalcohol mediated asymmetric addition of dimethylzinc to methylacrylaldehyde²⁹ (Scheme 5.2)

Methylation of aldehydes has been investigated less frequently, because dimethylzinc has a very low boiling point like diethyl ether. However, it can be handled easily as a toluene solution.



Scheme 5.2

Equipment

- Two-necked, round-bottomed flask (500 mL) fitted with rubber septum and a Teflon-coated magnetic stirring bar
- Three-way stopcock fitted to the top of the flask and connected to a vacuum/argon (or dry nitrogen) source
- Magnetic stirrer
- All-glass syringe with a needle-lock Luer and medium-gauge needle
- Water-jacketed short-path distillation apparatus
- Separating funnel (500 mL).

Protocol 2. Continued**Materials**

- Dry ether and hexane
- Methylacrylaldehyde
- Dimethylzinc
- (*R*)-1-*t*-Butyl-2-piperidinoethanol
- Technical ether for extraction

explosive on contact with air**flammable**

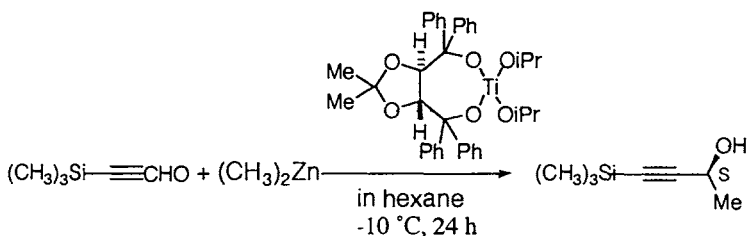
1. Flame dry the reaction vessel and place with a stirrer bar *in vacuo*. Fill the flask with argon and maintain a slightly positive argon pressure throughout the reaction.
2. Assemble the syringe and needle while hot and flush the syringe with argon.
3. Support the flask using clamp and a stand with a heavy base.
4. Charge the flask with hexane (100 mL) and there (100 mL) using a syringe by piercing the septum on reaction flask.
5. Keep the flask at -40°C with a dry ice/acetone bath.
6. Charge methylacrylaldehyde (6.6 g, 94.2 mmol) and (*R*)-1-*t*-butyl-2-piperidinoethanol (0.37 g, 2.0 mmol) with syringes under stirring.
7. Remove dimethylzinc in a syringe from the cylinder of dimethylzinc stored at -10°C under argon pressure.
8. Add dimethylzinc (10.5 mL, 102 mmol) dropwise from the syringe through the septum on the reaction flask to the mixture in the flask with vigorous stirring.
9. Allow the temperature of the flask to rise from -40°C to 15°C and keep the mixture at this temperature for 96 h.
10. Quench the mixture by adding 2 M HCl (100 mL) dropwise with stirring at 0°C .
11. Transfer the mixture to a separating funnel and separate into two layers. Extract the water layer with ether (2×50 mL).
12. Transfer the organic layer to a 500 mL flask. Dry the organic layer over Na_2SO_4 , and filter through a filter paper. Concentrate the filtrate by distillation under ordinal pressure.
13. Transfer the oily residue to a water-jacketed, short path distillation apparatus equipped with a thermometer. Distill the crude product under reduced pressure to obtain 3-methylbut-3-en-2-ol (b.p. $106\text{--}108^{\circ}\text{C}/760$ mm Hg; 6.8 g, 70% yield) as a colourless oil. The product is .99% pure by $^1\text{H-NMR}$; $\text{d}(\text{CDCl}_3)$ 1.29 (3H,d, J 6.7Hz), 1.60(1H,s), 1.75(3H,s), 4.25(1H,q, J 6.7Hz), 4.80(1H,s), and 4.96(1H,s).
14. Determine the absolute configuration of the product from the sign of $[\alpha]_D$. $[\alpha]_D -5.6$ (c 8.0, CHCl_3) showed it to be (*S*)-configurational product.

15. Enantiomeric excess(ee) (>95%) was determined by HPLC analysis of the 3,5-dinitrophenyl carbamate^a (column, Sumitomo Chemical Co. SUMIPAX OA 4000) in hexane/ethanol 100:1.5; t_R of (*R*)-isomer, 41 min; t_R of (*S*)-isomer, 36 min.

^a React 3-methylbut-3-en-2-ol (8.6 mg, 1.0 mmol) with 3,5-dinitrophenylisocyanide (21 mg, 1.0 mmol) in toluene (2.0 mL) at room temperature for 1 h with stirring. This mixture is used directly for HPLC.

Protocol 3.

Chiral Taddolated catalysed asymmetric addition of dimethylzinc to trimethylsilylpropynal¹⁰⁵ (Scheme 5.3)



Scheme 5.3

Equipment

- Two-necked, round-bottomed flask (500 mL) fitted with rubber septum and a teflon-coated magnetic stirring bar
- Three-way stopcock fitted to the top of the flask and connected to a vacuum/argon (or dry nitrogen) source
- Magnetic stirrer
- All-glass syringe with a needle-lock Luer and medium-gauge needle
- Water-jacketed short-path distillation apparatus
- Separatory funnel (500 mL)

Materials

- Dry toluene
- 3-Trimethylsilylpropynal
- Dimethylzinc
- (4*R*,5*R*)-2,2'-dimethyl- $\alpha,\alpha',\alpha',\alpha'$ -Tetraphenyl-1,3-dioxolane-4,5-dimethanol(TADDOL)
- Technical ether for extraction

explosive on contact with air

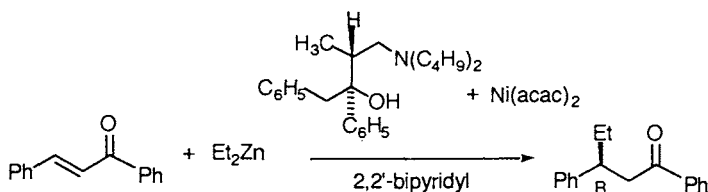
flammable

1. Flame dry the reaction vessel and place with a stirrer bar *in vacuo*. Fill the flask with argon and maintain a slightly positive argon pressure throughout the reaction.
2. Assemble the syringe and needle while hot and flush the syringe with argon.
3. Support the flask using clamp and a stand with a heavy base.

Protocol 3. Continued

4. Charge the flask with TADDOL (15.1 g, 20 mmol) and toluene (100 mL) using a syringe by piercing the septum on reaction flask.
5. Add titanium tetraisopropoxide (6 mL, 20 mmol) to the above mixture at room temperature and stir the resulting solution for 10 h.
6. Remove toluene and isopropanol produced under reduced pressure with stirring.
7. Dissolve the residue in *t*-butylmethyl ether (100 mL) solution containing titanium tetraisopropoxide (35 mL, 120 mmol).
8. Transfer dimethylzinc in a syringe from the cylinder of dimethylzinc stored at -10°C under an argon pressure.
9. Add dimethylzinc (12 mL, 180 mmol) dropwise from the syringe through the septum on the reaction flask to the mixture in the flask with vigorous stirring.
10. Add 3-trimethylsilylpropynal (16 g, 100 mmol) with a syringe with stirring at -20°C .
11. Allow the temperature of the flask to rise from -20°C to 0°C and keep the mixture at this temperature for 48 h.
12. Quench the mixture by adding 2M HCl(100 mL) dropwise with stirring at 0°C .
13. Transfer the mixture to a separating funnel and separate into two layers. Extract the water layer with ether (2×50 mL).
14. Transfer the organic layer to a 500 mL flask. Dry the organic layer over Na_2SO_4 , and filter through a filter paper. Concentrate the filtrate under reduced pressure using a rotary evaporator ($25^{\circ}\text{C}/50$ mm Hg).
15. Transfer the oily residue to a water-jacketed, short-path distillation apparatus equipped with a thermometer. Distil the crude product under reduced pressure to obtain 4-trimethylsilyl-3-butyn-2-ol (b.p. $77^{\circ}\text{C}/20$ mm Hg; 14.7 g, 98% yield) as a colourless oil. The product is 99% pure by ^1H NMR; $\delta(\text{CDCl}_3)$ 0.15(9H,s), 1.42 (3H,d, J 4.6 Hz), 1.85 (1H, d, J 3.8 Hz), and 4.50 (1H, m).
16. Determine the absolute configuration of the product from the sign of $[\alpha]_{\text{D}}$. $[\alpha]_{\text{D}} -24.2$ (c 2.0, CHCl_3) showed it to be (*S*)-configurational product.
17. Enantiomeric excess(ee) (96%) was determined by HPLC analysis of the 3,5-dinitrobenzoate^a (column, Daicel, CHIRAPAK AD in volume ratio of hexane/2-propanol (100:1); t_{R} of (*R*)-isomer, 30 min; t_{R} of (*S*)-isomer, 40 min.

^aReact 3-trimethylsilylbut-3-yn-2-ol (14.2 mg, 0.1 mmol) with 3,5-dinitrobenzoyl chloride (23 mg, 0.1 mmol) in toluene (2.0 mL) in the presence of triethylamine (1 mL) at room temperature for 1 h with stirring. Remove Et_3NHCl by filtration and use the filtrate directly for HPLC.

Protocol 4.**Asymmetric 1,4-conjugate addition of diethylzinc to chalcone catalysed by chiral nickel complex¹⁰⁶ (Scheme 5.4)****Scheme 5.4****Equipment**

- Two-necked, round-bottomed flask (300 mL) fitted with a rubber septum and a Teflon-coated magnetic stirring bar
- Three-way stopcock fitted to the top of flask and connected to a vacuum/argon (or dry nitrogen) source
- Magnetic stirrer
- All-glass syringe with a needle-lock Luer and medium-gauge needle
- Water-jacketed short-path distillation apparatus
- Separating funnel (300 mL)
- Rotary evaporator
- Chiralcel OK for HPLC

Materials

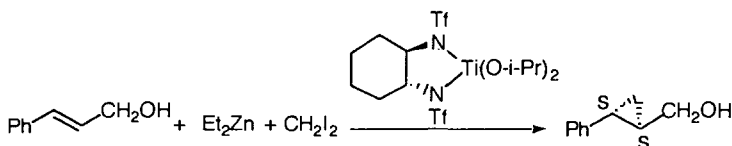
- Dry toluene
- Bis(acetylacetonato)nickel(II)
- *R*-(–)-(1*S*,2*R*)-*N,N*-Dibutylnorephedrine (DBNE)
- 2,2'-Bipyridyl
- Chalcone
- Diethylzinc
- Acetonitrile
- Hexane
- Technical ether for extraction

explosive on exposure to air**flammable**

1. Flame dry the reaction vessel and place with a stirrer bar *in vacuo*. Fill the flask with argon and maintain a slightly positive argon pressure throughout this reaction.
2. Assemble the syringe and needle while hot and flush the syringe with argon.
3. Support the flask using clamp and a stand with a heavy base.
4. Place bis(acetylacetonato)nickel(II) (3.6 g, 14 mmol) in the flask.
5. Add (1*S*,2*R*)-(–)-DBNE (0.896 g, 34 mmol) to the flask in CH₃CN (20 mL) using a syringe by piercing the septum on reaction flask. Stir the mixture at 80°C for 1 h.
6. Remove the acetylacetonate liberated and CH₃CN under reduced pressure.
7. Add 2,2'-bipyridyl (0.22 g, 14 mmol) in CH₃CN (20 mL) to the flask and stir at 80°C for 1 h.

Protocol 4. Continued

8. Cool the resulting green solution to room temperature.
9. Charge chalcone (4.16 g, 0.2 mol) in CH_3CN (100 mL) to above mixture followed by stirring for 20 min at room temperature.
10. Cool to -30°C this solution and add diethylzinc (1.0M toluene solution, 240 mL, 0.24 mol) dropwise from the syringe through the septum on the reaction flask and stir for 12 h at this temperature.
11. Quench the mixture by dropping 1M HCl (100 mL) with stirring at 0°C .
12. Transfer the mixture to a separating funnel and allow the two layers to separate. Extract the aqueous layer with ether (4×20 mL), and dry over anhydrous Na_2SO_4 .
13. Transfer the organic layer to a 300 mL round-bottomed flask and concentrate under reduced pressure using a rotary evaporator ($30^\circ\text{C}/30$ mm Hg).
14. Purify the residue by silica gel chromatography (eluent, hexane/chloroform, 1:1, v/v) to obtain (*R*)-(-)-1,3-diphenylpentan-1-one (2.26 g, 47% yield), which displays the appropriate ^1H NMR in CDCl_3 .
15. Determine the absolute configuration and enantiomeric excess (ee) of the product. The absolute configuration is determined from the sign of $[\alpha]_D$. $[\alpha]_D -4.7$ (c 2.5, EtOH) indicated (*R*)-configuration. Ee (90%) was determined by HPLC analysis (column, Daicel CHIRALCEL OD; hexane/2-propanol, 100:0.2, flow rate 0.5 mL/min; t_R of (*S*)-isomer, 40.1 min; t_R of (*R*)-isomer, 44.8 min).

Protocol 5.**Chiral disulfonamide-titanate catalysed asymmetric cyclopropanation of cinnamyl alcohol¹²⁴ (Scheme 5.5)****Scheme 5.5****Equipment**

- Two-necked, round-bottomed flask (500 mL) fitted with a rubber septum and a Teflon-coated magnetic stirring bar
- Three-way stopcock fitted to the top of the flask and connected to a vacuum/argon (or dry nitrogen) source
- Magnetic stirrer
- All-glass syringe with a needle-lock Luer and medium-gauge needle
- Water-jacketed short-path distillation apparatus
- Separating funnel (500 mL)

Materials

- Dry dichloromethane
- Di-iodomethane
- Diethylzinc
- (4*R*,5*R*)-2,2'-Dimethyl- $\alpha,\alpha,\alpha',\alpha'$ -tetraphenyl-1,3-dioxolane-4,5-dimethanol(TADDOL)
- Ti (OiPr)₄
- Technical ether for extraction
- Molecular sieve 4A and Ti(OiPr)₄

explosive on exposure to air**flammable****flammable**

1. Flame dry the reaction vessel and place with a stirrer bar *in vacuo*. Fill the flask with argon and maintain a slightly positive argon pressure throughout the section.
2. Assemble the syringe and needle while hot and flush the syringe with argon.
3. Support the flask using clamp and a stand with a heavy base.
4. Charge the flask with CH₂I₂ (1.6 mL, 20 mmol) and CH₂Cl₂ (50 mL) using a syringe by piercing the septum on reaction flask.
5. Remove diethylzinc in a syringe from the cylinder of diethylzinc stored under argon pressure.
6. Add diethylzinc (1 mL, 10 mmol) dropwise to the flask at 0°C and stir the resulting solution for 15 min. A white precipitate is formed (solution **A**).
7. Mix (4*R*,5*R*)-TADDOL (1.4 g, 2.9 mmol) and molecular sieve 4A (10 g) in CH₂Cl₂ (5 mL), and add Ti(OiPr)₄ (0.74 mL, 2.5 mmol) with stirring. Stir for 1.5 h at room temperature.
8. Remove the solvent and isopropanol produced by the reaction under reduced pressure and leave the residue under high vacuum. Dissolve the residue into CH₂Cl₂ (50 mL) (solution **B**).
9. Mix solutions **A** and **B** in the flask at -40°C and add immediately cinnamyl alcohol (1.4 g, 10.4 mmol) to the mixture with stirring.
10. Stir the mixture at 0°C for 90 min and add TiCl₄ (0.16 mL, 1.5 mmol).
11. Cool the solution to -40°C and quench the reaction by pouring the mixture into saturated aq. NH₄Cl solution (300 mL).
12. Transfer the mixture to a separating funnel and allow the two layers to separate. Extract the water layer with ethyl acetate (50 mL).
13. Wash the organic layer with saturated aqueous NH₄Cl, and brine, then dry over MgSO₄ and concentrate under reduced pressure (25°C/30 mm Hg).
14. Purify the crude product by flash chromatography (EtOAc/hexane, 80:20) to obtain (1.2 g, 80% yield) as a colourless oil. The product is .99% pure by ¹H NMR.
15. Determine the absolute configuration of the product from the sign of [α]_D. [α]_D + 84 (c 1.3, EtOH) indicated a (2*SS*,3*SS*)- configurational product.
16. Enantiomeric excess(ee) (90%) was determined by GC analysis of the tri-fluoroacetate derivative on a chiral stationary phase: Cyclodex GT-A column, 0.32 mm × 30 m (25 psi, 110°C), t_r 11.5 min (minor), 12.0 (major).

References

1. Noyori, R.; Kitamura, M. *Angew. Chem. Int. Ed.* **1991**, *30*, 49–69.
2. Soai, K.; Niwa, S. *Chem. Rev.* **1992**, *92*, 833–856.
3. Knochel, P.; Singer, R. D. *Chem. Rev.* **1993**, *93*, 2117–2188.
4. Knochel, P. *Synlett* **1995**, 393–403.
5. Oguni, N.; Omi, T. *Tetrahedron Lett.* **1984**, *25*, 2823–2824.
6. Kitamura, M.; Suga, S.; Kawai, K.; Noyori, R. *J. Am. Chem. Soc.* **1986**, *108*, 6071–6072.
7. Alberts, A. H.; Wynberg, H. *J. Am. Chem. Soc.* **1989**, *111*, 7265–7266.
8. Noyori, R.; Suga, S.; Kawai, K.; Okada, S.; Kitamura, M.; Oguni, N.; Hayashi, M.; Kaneko, K.; Matsuda, Y. *J. Organometal. Chem.* **1990**, *382*, 19–37.
9. Smaardijk, A. A.; Wynberg, H. *J. Org. Chem.* **1987**, *52*, 135–137.
10. Soai, K.; Ookawa, A.; Ogawa, K.; Kaba, T. *J. Chem. Soc. Chem. Commun.* **1987**, 467–468.
11. Soai, K.; Ookawa, A.; Kaba, K.; Ogawa, K. *J. Am. Chem. Soc.* **1987**, *109*, 7111–7115.
12. Chaloner, P. A.; Perera, S. A. R. *Tetrahedron Lett.* **1987**, *28*, 3013–3016.
13. Soai, K.; Niwa, S.; Yamada, Y.; Inoue, H. *Tetrahedron Lett.* **1987**, *28*, 4841–4842.
14. Soai, K.; Yokoyama, S.; Ebihara, K.; Hayasaka, T. *J. Chem. Soc. Chem. Commun.* **1987**, 1690–1691.
15. Soai, K.; Nishi, S.; Ito, Y. *Chem. Lett.* **1987**, 2405–2406.
16. Corey, E. J.; Hannon, F. J. *Tetrahedron Lett.* **1987**, *28*, 5233–5236; 5237–5240.
17. Muchow, G.; Vannonoorenberghe, Y.; Buono, G. *Tetrahedron Lett.* **1987**, *28*, 6163–6166.
18. Soai, K.; Niwa, S.; Watanabe, M. *J. Org. Chem.* **1988**, *55*, 927–928.
19. Soai, K.; Watanabe, M.; Koyano, M. *J. Chem. Soc. Chem. Commun.* **1989**, 534–536.
20. Corey, E. J.; Yuen, P.; Hannon, F. J.; Wierda, D. A. *J. Org. Chem.* **1990**, *55*, 784–746.
21. Oppolzer, W.; Radinov, R. N. *Tetrahedron Lett.* **1988**, *29*, 5645–5648.
22. Itsuno, S.; Frechet, J. M. J. *J. Org. Chem.* **1987**, *52*, 4140–4142.
23. Itsuno, S.; Sakurai, Y.; Ito, K.; Maruyama, T.; Nakahama, S.; Frechet, J. M. J. *J. Org. Chem.* **1990**, *55*, 304–310.
24. Tanaka, K.; Ushio, H.; Suzuki, H. *J. Chem. Soc. Chem. Commun.* **1989**, 1700–1701.
25. van Oeveren, A.; Menge, W.; Feringa, B. L. *Tetrahedron Lett.* **1989**, *30*, 6427–6430.
26. Soai, K.; Watanabe, M. *J. Chem. Soc., Chem. Commun.* **1990**, 43–44.
27. Chaloner, P. A.; Longadianon, E. *Tetrahedron Lett.* **1990**, *31*, 5185–5186.
28. Niwa, S.; Soai, K. *J. Chem. Soc., Perkin Trans I* **1990**, 937–943.
29. Hayashi, M.; Kaneko, T.; Oguni, N. *J. Chem. Soc. Perkin. Trans I* **1991**, 25–28.
30. Rosini, C.; Franzini, L.; Iuliano, A.; Pini, D.; Salvadori, P. *Tetrahedron Asymmetry* **1990**, *1*, 587–588; **1991**, *2*, 363–366.
31. Oppolzer, W.; Radinov, R. N. *Tetrahedron Lett.* **1991**, *32*, 5777–5780.
32. Hayashi, M.; Miwata, H.; Oguni, N. *Chem. Lett.* **1989**, 1969–1970.
33. Hayashi, M.; Miwata, H.; Oguni, N. *J. Chem. Soc. Perkin Trans I* **1991**, 1167–1171.
34. Soai, K.; Niwa, S.; Hatanaka, T. *J. Chem. Soc., Chem. Commun.* **1990**, 709–711.

35. Hatanaka, T.; Yamashita, T.; Soai, K. *J. Chem. Soc., Chem. Commun.* **1992**, 927–928.
36. Conti, S.; Falorni, M.; Giacomelli, G.; Soccolini, F. *Tetrahedron* **1992**, *48*, 8993–9000.
37. Matsumoto, Y.; Ohno, A.; Lu, S.; Hayashi, T.; Oguni, N.; Hayashi, M. *Tetrahedron Asymmetry* **1993**, *4*, 1763–1766.
38. Vries, E. F. J.; Brussee, J.; Kruse, C. G.; van der Gen, A. *Tetrahedron Asymmetry* **1993**, *4*, 1987–1990.
39. Watanabe, M.; Soai, K. *J. Chem. Soc., Perkin Trans I* **1994**, 837–842, 3125–3128.
40. Soai, K.; Shimada, C.; Takeuchi, M.; Itabashi, M. *J. Chem. Soc. Chem. Commun.* **1994**, 567–568.
41. Soai, K.; Hayase, T.; Takai, K.; Sugiyama, T. *J. Org. Chem.* **1994**, *59*, 7908–7909.
42. Shi, M.; Satoh, Y.; Makihara, T.; Masaki, Y. *Tetrahedron Asymmetry* **1995**, *6*, 2109–2112.
43. Dai, W. M.; Zhu, H. J.; Hao, X.-J. *Tetrahedron Asymmetry* **1995**, *6*, 1857–1860.
44. Nakano, H.; Kumagai, H.; Kabuto, C.; Matsuzaki, H.; Hongo, H. *Tetrahedron Asymmetry* **1995**, *6*, 1233–1236.
45. Collomb, P.; von Zelewsky, A. *Tetrahedron Asymmetry* **1995**, *6*, 2903–2904.
46. Soai, K.; Inoue, Y.; Takahashi, T.; Shibata, T. *Tetrahedron* **1996**, *52*, 13355–13362.
47. Peper, V.; Martens, J. *Chem. Ber.* **1996**, *129*, 691–698.
48. Audres, J. M.; Barrio, R.; Martine, M. A.; Redrosa, R. *J. Org. Chem.* **1996**, *61*, 4210–4213.
49. Falorni, M.; Collu, C.; Conti, S.; Giacomelli, G. *Tetrahedron Asymmetry* **1996**, *7*, 293–299.
50. Kitajima, H.; Ito, K.; Katsuki, T. *Chem. Lett.* **1996**, 343–344.
51. Bolm, C.; Zehnder, M.; Bur, D. *Angew. Chem. Int. Ed.* **1990**, *29*, 205–207.
52. Bolm, C.; Schlingloff, G.; Harmo, K. *Chem. Ber.* **1992**, *125*, 1191.
53. Ishizuka, M.; Hoshino, O. *Chem. Lett.* **1994**, 1337–1338.
54. Ishizaki, M.; Hoshino, O. *Tetrahedron Asymmetry* **1994**, *5*, 1901–1904.
55. Ishizaki, M.; Fujita, K.; Shimamoto, M.; Hoshino, O. *Tetrahedron Asymmetry* **1994**, *5*, 411–412.
56. Macedo, E.; Moberg, C. *Tetrahedron Asymmetry* **1995**, *6*, 549–548.
57. Kotsuki, H.; Hayakawa, H.; Wako, M.; Shimanouchi, T.; Ochi, M. *Tetrahedron*.
58. Wally, H.; Widhalm, M.; Weisensteiner, W.; Schlogl, K. *Tetrahedron Asymmetry* **1993**, *4*, 285–288.
59. Watanabe, M.; Komota, M.; Nishimura, M.; Araki, S.; Butsugan, Y. *J. Chem. Soc. Perkin Trans I* **1993**, *18*, 2193–2196.
60. Nicolosi, G.; Patti, A.; Morrone, R.; Piattelli, M. *Tetrahedron Asymmetry* **1994**, *5*, 1639–1642.
61. Watanabe, M. *Synlett* **1995**, 1050–1052.
62. Watanabe, M. *Tetrahedron Lett.* **1995**, *36*, 8991–8994.
63. Fukuzawa, S.; Tsuduki, K.; *Tetrahedron Asymmetry* **1995**, *6*, 1039–1942.
64. Watanabe, M.; Araki, S.; Butsugan, Y.; Uemura, M. *J. Org. Chem.* **1991**, *56*, 2218–2224.
65. Uemura, M.; Miyake, R.; Nakayama, K.; Shiro, M.; Hayashi, Y. *J. Org. Chem.* **1993**, *58*, 1238–1244.
66. Heaton, S. B.; Jones, G. B. *Tetrahedron Lett.* **1992**, *33*, 1693–1696.
67. Jones, G. B.; Heaton, S. B. *Tetrahedron Asymmetry* **1993**, *4*, 261–262.

68. Rijnberg, E.; Jastrzebski, J. T. B. H.; Janssen, M. D.; Boersma, J.; Koten, G. *Tetrahedron Lett.* **1994**, 35, 6521–6524.
69. Kang, J.; Kim, D. S.; Kim, J. I. *Synlett* **1994**, 842–843.
70. Kang, J.; Lee, J. W.; Kim, J. I. *J. Chem. Soc., Chem. Commun.* **1994**, 2009–2010.
71. Fitzpatrick, K.; Hulst, R.; Kellogg, R. M. *Tetrahedron Asymmetry* **1995**, 6, 1861–1864.
72. Hulst, R.; Heres, H.; Fitzpatrick, K.; Peper, N. C. M. W.; Kellogg, R. M. *Tetrahedron Asymmetry* **1996**, 7, 2755–2760.
73. Oguni, N.; Matsuda, Y.; Kaneko, T. *J. Am. Chem. Soc.* **1988**, 110, 7877–7878.
74. Kitamura, M.; Okada, S.; Suga, S.; Noyori, R. *J. Am. Chem. Soc.* **1989**, 111, 4028–4036.
75. Kitamura, M.; Suga, S.; Niwa, M.; Noyori, R. *J. Am. Chem. Soc.* **1995**, 117, 4832–4842.
76. Kitamura, M.; Yamakawa, M.; Oka, da, S.; Suga, S.; Noyori, R. *Chem. Eur. J.* **1996**, 2, 1173–1181.
77. Wynberg, H. *Chimia* **1989**, 43, 150–152.
78. Soai, K.; Hayase, T.; Takai, K.; Sugiyama, T. *J. Org. Chem.* **1994**, 59, 7908–7909.
79. Soai, K.; Hayase, T.; Shimada, C.; Isobe, K. *Tetrahedron Asymmetry* **1994**, 5, 789–792.
80. Soai, K.; Hayase, T.; Takai, K. *Tetrahedron Asymmetry* **1995**, 6, 637–638.
81. Soai, K.; Shibata, T.; Morioka, H.; Choji, K. *Nature* **1995**, 378, 767–768.
82. Shibata, T.; Morioka, H.; Hayase, T.; Choji, K.; Soai, K. *J. Am. Chem. Soc.* **1996**, 118, 471–472.
83. Oguni, N.; Omi, T.; Yamamoto, Y.; Nakamura, A. *Chem. Lett.* **1983**, 841–842.
84. Yoshioka, M.; Kawakita, T.; Ohno, M. *Tetrahedron Lett.* **1989**, 30, 1657–1660.
85. Takahashi, H.; Kawakita, T.; Yoshioka, M.; Kobayashi, S.; Ohno, M. *Tetrahedron Lett.* **1989**, 30, 7095–7098.
86. Joshi, N. N.; Screbnik, M.; Brown, H. C. *Tetrahedron Lett.* **1989**, 30, 5551–5554.
87. Schmidt, B.; Seebach, D. *Angew. Chem.* **1991**, 103, 100–101.
88. Schmidt, B.; Seebach, D. *Angew. Chem.* **1991**, 30, 99–101.
89. Seebach, D.; Behrendt, L.; Felix, D. *Angew. Chem.* **1991**, 30, 1008–1009.
90. Schmidt, B.; Seebach, D. *Angew. Chem.* **1991**, 30, 1231–1233.
91. Bussche-Hunnefeld, J. L.; Seebach, D. *Tetrahedron* **1992**, 48, 5719–5730.
92. Itoh, K.; Kimura, Y.; Okamura, H.; Katsuki, T. *Synlett* **1992**, 573–574.
93. Mori, A.; Yu, D.; Inoue, S. *Synlett* **1992**, 427–428.
94. Knochel, P.; Brieden, W.; Rozema, M. J.; Eisenberg, C. *Tetrahedron Lett.* **1993**, 34, 5881–5884.
95. Soai, K.; Hirose, Y.; Ohno, Y. *Tetrahedron Asymmetry* **1993**, 4, 1473–1474.
96. Klement, I.; Knochel, P. *Tetrahedron Lett.* **1994**, 35, 1177–1180.
97. Schwink, L.; Knochel, P. *Tetrahedron Lett.* **1994**, 35, 9007–9010.
98. Vettel, S.; Knochel, P. *Tetrahedron Lett.* **1994**, 34, 5849–5852.
99. Ostwald, R.; Chavant, P.-Y.; Stadtmüller, H.; Knochel, P. *J. Org. Chem.* **1994**, 59, 4143–4153.
100. Lutjens, H.; Knochel, P. *Tetrahedron Asymmetry* **1994**, 5, 1161–1162.
101. Nowotny, S.; Vettel, S.; Knochel, P. *Tetrahedron Lett.* **1994**, 35, 4539–4640.
102. Seebach, D.; Beck, A. K.; Schmidt, B.; Wang, Y. M. *Tetrahedron* **1994**, 50, 4363–4384.

103. Lutjens, H.; Nowotny, S.; Knochel, P. *Tetrahedron Asymmetry* **1995**, 6, 2675–2678.
104. Vettel, S.; Vaupel, A.; Knochel, P. *Tetrahedron Lett.* **1995**, 36, 1023–1026.
105. Oguni, N.; Satoh, N.; Fujii, H. *Synlett* **1995**, 1043–1044.
106. Soai, K.; Hayasaka, T.; Ukaji, S. *J. Chem. Soc. Chem. Commun.* **1989**, 516–517.
107. Bolm, C. *Tetrahedron Asymmetry* **1991**, 2, 701–704.
108. Uemura, M. *Tetrahedron Asymmetry* **1992**, 3, 713–714.
109. Felder, M.; Muller, J.; Bolm, C. *Synlett* **1992**, 439–441.
110. Asami, M.; Usui, K.; Higuchi, S.; Inoue, S. *Chem. Lett.* **1994**, 297–298.
111. Kobayashi, S. *Tetrahedron Lett.* **1992**, 33, 2575–2578.
112. Denmark, S. E. *Synlett* **1992**, 229–230.
113. Ukaji, Y. *Chem. Lett.* **1992**, 61–64.
114. Ukaji, Y.; Sada, K.; Inomata, K. *Chem. Lett.* **1993**, 1227–1230.
115. Imai, N.; Sakamoto, K.; Takahashi, H.; Kobayashi, S. *Tetrahedron Lett.* **1994**, 35, 7045–7048.
116. Imai, N.; Takahashi, H.; Kobayashi, S. *Chem. Lett.* **1994**, 177–180.
117. Charette, A. B.; Brochu, C. *J. Am. Chem. Soc.* **1995**, 117, 11367–11368.
118. Charette, A. B.; Brochu, C. *J. Org. Chem.* **1995**, 60, 181–183.
119. Audres, J. M.; Charette, B.; Marcoux, J. F. *Synlett* **1995**, 1197–1207.
120. Takahashi, H.; Yoshioka, Y.; Shibasaki, M.; Ohno, M.; Imai, N.; Kobayashi, S. *Tetrahedron* **1995**, 51, 12013–12026.
121. Kitajima, H.; Aoki, Y.; Ito, K.; Katsuki, T. *Chem. Lett.* **1995**, 1113–1114.
122. Charette, A. B.; Prescott, S.; Brochu, C. *J. Org. Chem.* **1995**, 1081–1083.
123. Denmark, S. E.; Christensen, B. L.; Coe, D. M.; O’Cornor, S. P. *Tetrahedron Lett.* **1995**, 36, 2215–2218; 2219–2122.
124. Charette, A. B.; Brochu, C. *J. Am. Chem. Soc.* **1995**, 117, 11367–11368.
125. Ukaji, Y.; Sada, K.; Inomata, K. *Chem. Lett.* **1993**, 1847–1850.
126. Anderson, P. G.; Guijarro, D.; Tanner, D. *Synlett* **1996**, 727–728.
127. Hayase, T.; Inoue, Y.; Shibata, T.; Soai, K. *Tetrahedron Asymmetry* **1996**, 7, 2509–2510.
128. Suzuki, T.; Narisada, N.; Shibata, T.; Soai, K. *Tetrahedron Asymmetry* **1996**, 7, 2519–2522.
129. Nakamura, N.; Hirai, A.; Nakamura, E. *J. Am. Chem. Soc.* **1996**, 118, 8489–8490.

Chiral titanium complexes for enantioselective catalysis

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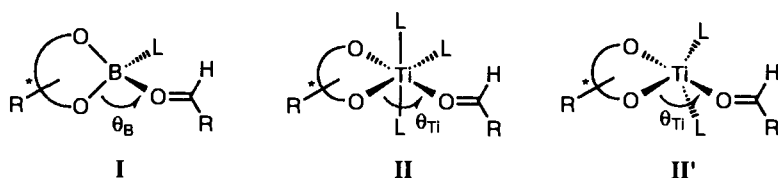
1. Introduction

Enantioselective catalysis is an economical and environmentally benign process, since it achieves 'multiplication of chirality'¹ thereby affording a large amount of the enantio-enriched product, while producing a small amount of waste material, due to the very small amount of chiral catalyst used. Thus, the development of enantioselective catalysts is a most challenging and formidable endeavour for synthetic organic chemists.^{2,3} Highly promising candidates for such enantioselective catalysts are metal complexes bearing chiral organic ligands. Thus, the choice of the central metal and the molecular design of the chiral organic ligand are most crucial for the development of enantioselective catalysts. The degree of enantioselectivity should be critically influenced by metal–ligand bond lengths, particularly metal–oxygen and –nitrogen bond lengths in case of metal alkoxide and amide complexes (Table 6.1),^{4,5} as well as the steric demand of the organic ligands. The shorter the bond length, in principle, the higher the enantioselectivity, because the asymmetric environment constructed by the chiral ligand is closer to the reaction centre. Therefore, boron and aluminium are the main group elements of choice, and

Table 6.1 Metal-oxygen (M-O) and metal-nitrogen (M-N) bond length

| Metal | M-O bond length (Å) | M-N bond length (Å) |
|-------|---------------------|---------------------|
| Li | 1.92–2.00 | 2.12 |
| Mg | 2.01–2.13 | 2.22 |
| Zn | 1.92–2.16 | 2.16 |
| Sn | 2.70 | 2.25 |
| Al | 1.92 | 1.95 |
| B | 1.36–1.47 | 1.40 |
| Ti | 1.62–1.73 | 2.30 |
| Zr | 2.15 | 2.30 |

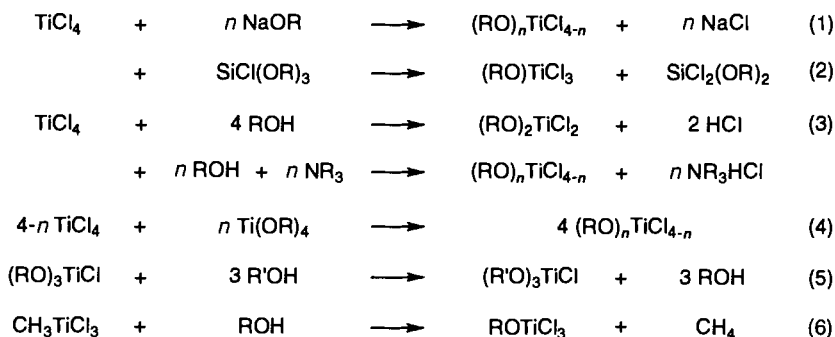
titanium is one of the best early transition metals with hexa- and penta-co-ordination (Scheme 6.1).



Scheme 6.1

The Lewis acidity of metal complexes is generally proportional to the value of (charge density) \times (ionic radius)^{-3.6}. Electron donating and sterically demanding ligands decrease the Lewis acidity but increase the configurational stability of titanium complexes in the order: Cp (cyclopentadienyl) > NR₂ > OR > X (halides).^{4,7,8} Therefore, the Lewis acidity of titanium complexes decreases on going from titanium halides to titanium alkoxides to titanium amides. The Lewis acidity can be fine-tuned by mixing ligands such as in alkoxytitaniumhalide complexes. By contrast, bond strengths with titanium decrease in the order: M–O > M–Cl > M–N > M–C.⁴ The M–C bond strengths in Ti(IV) compounds are comparable with those of other metal–carbon bonds. However, the Ti–O bond is exceptionally strong.

General preparative procedures for chiral titanium complexes are classified in Scheme 6.2.^{9–13} (1) A halide can be replaced with metalated ligands by transmetalation (eq. 1). (2) A halide can be replaced by metathesis of a silylated ligand with accompanying generation of silylhalide (eq. 2). (3) HCl is evolved with protic ligands, and hence must be evaporated or neutralized with a base (eq. 3). (4) Ligand redistribution results in disproportionation (eq. 4) (Protocol 1). (5) A chiral titanate ester is prepared using an alkoxy exchange reaction (*trans*-esterification) with a free chiral alcohol (eq. 5). The equilibrium is shifted towards the chiral titanium complex by azeotropic removal of



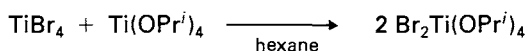
Scheme 6.2

the volatile achiral alcohol. (6) Alkyl (methyl, in particular) titanium complexes can be used for deprotonation of the chiral ligands along with generation of alkane (methane) (eq. 6). Chiral titanium alkoxide complexes thus obtained usually form bridged dimers, or trimers in extreme cases. Such aggregates are the favoured form even in solution. As shown above, there are many ways to prepare chiral titanium complexes for the enantioselective catalysis of carbon–carbon bond forming reactions or asymmetric functional group transformations. However, reactions of the latter type such as the Sharpless oxidation^{14–17} or reduction^{18–23} are beyond the scope of this manuscript due to limited space.

Protocol 1.

Preparation of di-isopropoxytitanium(IV) dibromide *via* ligand redistribution^{24,25}

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Equipment

- A pre-dried, pre-weighed, two-necked, round-bottomed flask (50 mL) equipped with a three-way stopcock, a magnetic stirring bar, and a septum cap. The three way stopcock is connected to a vacuum/argon source
- Dry gas-tight syringe
- Thermometer for low temperatures

Materials

- | | |
|--|---|
| • Titanium(IV) bromide (FW 367.5), 7.35 g, 20 mmol | corrosive, lachrymator, moisture sensitive |
| • Titanium(IV) isopropoxide (FW 284.3), 5.94 mL, 20 mmol | flammable liquid, irritant, moisture sensitive |
| • Dry hexane, total 40 mL | flammable liquid, irritant |
| • Dry toluene, 87.6 mL | flammable liquid, toxic |

1. Flame dry reaction vessel and accessories under dry argon. After cooling to room temperature, add the titanium (IV) bromide (7.35 g, 20 mmol) to the pre-weighed, two-necked, round-bottomed flask under a flow of argon.
2. Introduce 20 mL of hexane to this flask to form a red–brown suspension and then add the titanium(IV) isopropoxide (5.94 mL, 20 mmol) carefully over ~7 min at ambient temperature with the syringe. The addition of titanium (IV) isopropoxide causes the mixture to warm to about 40°C.
3. Stir for 10 min at that temperature to give a yellow solution. Then allow to stand for 6 h at room temperature. Isolate the pale-yellow precipitate that has formed by removal of the supernatant solvent with a syringe.

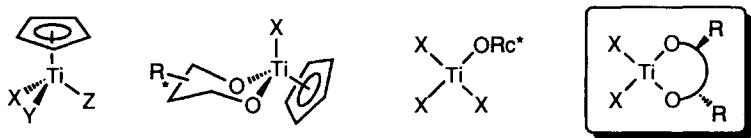
Protocol 1. Continued

4. Wash with hexane (5 mL \times 2) and recrystallize from hexane (10 mL). The recrystallization is carried out in the same flask by heating to reflux and then leaving the solution at room temperature overnight.
5. Remove the supernatant solvent again with the syringe. Vacuum dry the pale-yellow crystalline solid to give di-isopropoxytitanium(IV) dibromide (5.71 g, 44%), a highly moisture sensitive product.
6. Dissolve the crystalline in dry toluene (87.6 mL) to give a 0.2 M solution. Store the solution in a refrigerator.

Additionally, the crystalline solid may be stored at 0 \sim -20°C .

2. Carbonyl addition reaction

Alkyltitanium complexes can be obtained from alkaline metal carbanions via titanation. Introduction of chirality at the titanium centre or on the ligand (or a combination of both) (Scheme 6.3) allows for the possibility of asymmetric induction in the carbonyl addition reaction.



Scheme 6.3

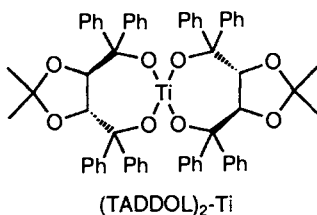
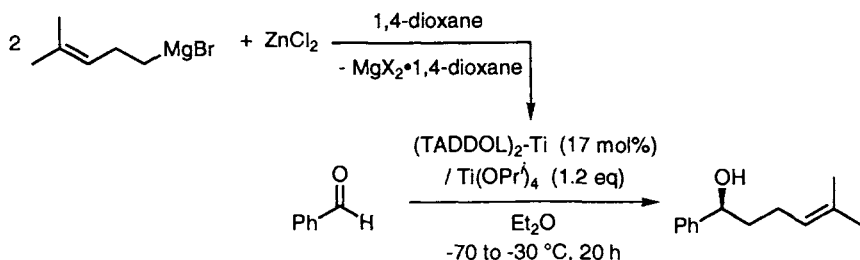
However, the use of complexes which are chiral at the titanium centre, which is the closest to the reacting carbonyl group, generally affords only low enantioselectivity, because of the configurational lability of chiral titanium centre.²⁶ The use of a C_2 symmetric 1,1'-bi-2-naphthol (BINOL)²⁷ derived titanium complex has been unsuccessful so far in allylation or methylation reactions.⁴ In an exceptional case, high enantioselectivity was obtained with a BINOL-modified phenyltitanium reagent. In that reaction, chiral titanium 'ate' complexes formed from BINOL-Ti(OPr^{*i*})₂²⁸ and arylmagnesium halides could also be used.^{29,30} Allylation of aromatic and aliphatic aldehydes by (4*R*,5*R*)- $\alpha,\alpha,\alpha',\alpha'$ -tetra-aryl-1,3-dioxolane-4,5-dimethanol (TADDOL)^{31–34}-derived cyclopentadienyltitanium complexes was found to give homoallyl alcohols with high enantioselectivity.³⁵ Exploitation of the Schlenk equilibrium of a mixture of a Grignard reagent, RMgX and 0.5 equiv. of ZnCl₂ with dioxane allowed the *in situ* generation of a functionalized alkyl zinc reagent along with MgX₂-dioxane complex (Protocol 2).^{36,37} Such R₂Zn can be used as nucleophile for (TADDOL)₂-Ti complex-catalysed carbonyl addition reaction with

high enantioselectivity. Chiral titanium bis-sulfonylamine complexes have also been used to accelerate the addition reaction of dialkylzinc.^{38–43}

Protocol 2.

TADDOL–Ti complex-catalysed asymmetric carbonyl addition reaction with functionalized dialkylzinc compounds, generated *in situ* from Grignard reagents^{36,37} (Scheme 6.4)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 6.4

Equipment

- A pre-dried, three-necked round-bottomed flask (100 mL) equipped with a condenser, a three-way stopcock, a magnetic stirring bar, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- A pre-dried, three-necked, round-bottomed flask (100 mL) equipped with a magnetic stirring bar, a glass filter, a three-way stopcock, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- A pre-dried, two-necked, round-bottomed flask (100 mL) equipped with a three-way stopcock, a magnetic stirring bar, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- Dry gas-tight syringe
- Thermometers for low temperature

Materials

- (4*R*,5*R*)-2,2-dimethyl- $\alpha,\alpha,\alpha',\alpha'$ -tetraphenyl-1,3-dioxolane-4,5-dimethanol (FW 466.6), 9.33 g, 20 mmol
- Distilled titanium(IV) ethoxide (FW 228.2), 2.3 mL, 11 mmol
- Dry toluene, 20 mL

flammable liquid, moisture sensitive
flammable liquid, toxic

Protocol 2. Continued

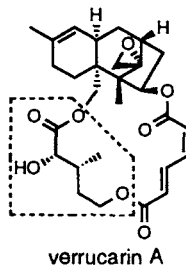
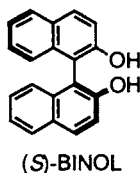
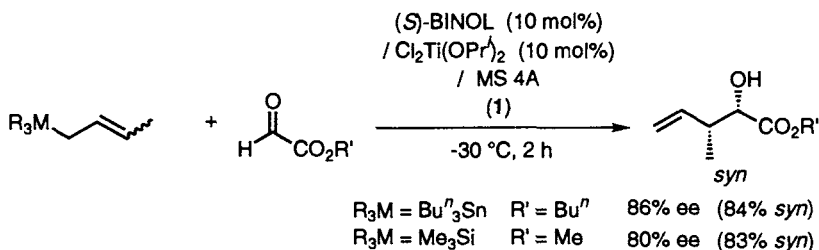
- Titanium(IV) isopropoxide (FW 284.3), 1.80 mL, 6 mmol **flammable liquid, irritant, moisture sensitive**
 - 1.0 M solution of zinc chloride in ether, 10 mL, 10 mmol **flammable liquid, moisture sensitive**
 - 1.05 M solution of 4-methyl-3-penten-1-ylmagnesium bromide in ether, 19 mL, 20 mmol **flammable liquid, moisture sensitive**
 - Dry ether **flammable liquid, toxic**
 - Dry 1,4-dioxane (FW 88.1), 6 mL, 70 mmol **cancer suspect agent, flammable liquid**
 - Benzaldehyde (FW 106.1) 0.51 mL, 5 mmol **highly toxic, cancer suspect agent**
1. Add the (4*R*,5*R*)-2,2-dimethyl- $\alpha,\alpha,\alpha',\alpha'$ -tetraphenyl-1,3-dioxolane-4,5-dimethanol (9.33 g, 20 mmol) to the pre-dried three-necked round-bottomed flask with condenser, and purge with argon.
 2. Add the dry toluene (20 mL) and the distilled titanium(IV) ethoxide (2.3 mL, 11 mmol), giving to rise a slightly yellow suspension.
 3. Stir for 12 h at 40°C. Heat the resulting clear solution to reflux temperature for an additional 5 h.
 4. Slowly evaporate the solvent *in vacuo*. Control the speed of the distillation by regulation of the vacuum.
 5. Dry the resulting yellow, waxy solid *in vacuo* to obtain the (TADDOL)₂-Ti complex in quantitative yield. The (TADDOL)₂-Ti complex should be stored under argon.
 6. Add the (TADDOL)₂-Ti complex (0.856 g, 0.85 mmol) to the other pre-dried three-necked round-bottomed flask fitted with a glass filter. Dissolve the (TADDOL)₂-Ti complex with ether (12 mL) and introduce the titanium(IV) isopropoxide (1.8 mL, 6 mmol) by the syringe.
 7. Introduce the 1.0 M ether solution of zinc chloride (10 mL, 10 mmol) into the pre-dried two-necked round bottomed flask with a syringe and dilute with ether (5 mL).
 8. Add 1.05 M Grignard reagents (19 mL, 20 mmol), 4-methyl-3-penten-1-ylmagnesium bromide,^a and stir at room temperature for 2 h.
 9. Add 1,4-dioxane (6 mL, 70 mmol) to the suspension and stir for an additional 45 min. To obtain a clear solution of the dialkyl zinc reagent, filter under an argon atmosphere. Add the filtrate directly to the TADDOL-Ti solution, obtained above, at -78°C.
 10. After stirring at -78°C for 1 h, add benzaldehyde (0.51 mL, 5 mmol) and then raise the reaction temperature to -30°C.
 11. Stir at -30°C for 20 h and then quench with saturated NH₄Cl solution (10 mL) and add ether (25 mL) at -30°C. Filter through a pad of Celite and rinse the filter cake with ether. Wash the organic phase with H₂O and brine, dry over Na₂SO₄, filter the solution, and concentrate *in vacuo*.
 12. To crystallize and separate from the TADDOL ligand, add pentane to the

resulting crude product. Again concentrate the supernatant solution and isolate by bulb to bulb distillation to collect 0.847 g (89%) of 1-phenyl-5-methylhex-4-en-1-ol.

13. The enantiomeric purity is determined by the capillary GC analysis using a heptakis(2,3,6-tri-*O*-ethyl)- β -cyclodextrin in OV 1701 as a chiral stationary phase column (96% ee).

^a Prepared as usual from 5-bromo-2-methylpent-2-ene and magnesium turnings and titrated shortly before use.

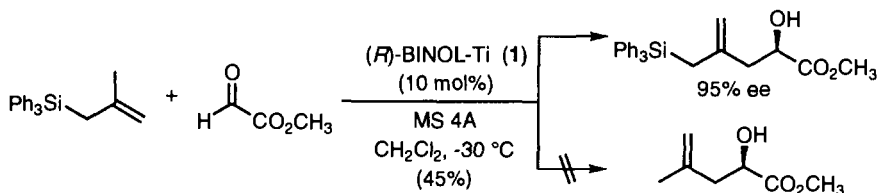
Further pursuing our research project on the asymmetric catalysis of the carbonyl-ene reaction, we have found that the BINOL-Ti complexes (**1**),⁴⁴ which are prepared *in situ* in the presence of MS 4A from di-isopropoxy-titanium dihalides ($X_2Ti(OPr^i)_2$; $X = Br$ or Cl) and optically pure BINOL (see below), catalyse rather than promote stoichiometrically the carbonyl addition reaction of allylic silanes and stannanes.⁴⁵ The addition reactions to glyoxylate of (*E*)-2-butenylsilane and -stannane proceed smoothly to afford the *syn*-product in high enantiomeric excess (Scheme 6.5). The *syn*-product thus obtained could readily be converted into the lactone portion of verrucarline A.^{46,47}



Scheme 6.5

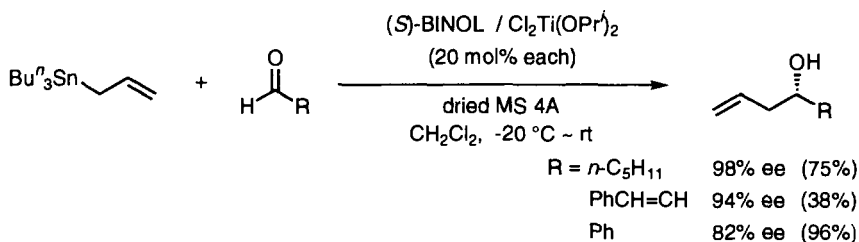
We have further found the BINOL-Ti (**1**) catalysis for the Sakurai-Hosomi reaction of methallylsilanes with glyoxylates (Scheme 6.6).⁴⁸ Surprisingly,

however, the products were obtained in the allylic silane (ene product) form (see below), with high enantioselectivity.

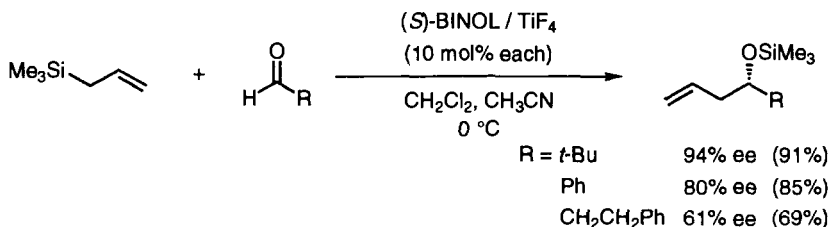


Scheme 6.6

Asymmetric catalysis by BINOL–Ti complexes of the reaction of aliphatic and aromatic aldehydes with an allylstannane has been reported independently by Tagliavini, Umami-Ronchi, and Keck.^{49–52} In Tagliavini's case,⁴⁹ a new complex generated by reaction of the BINOL–Ti complex with allylstannane has been suggested to be the catalytic species which provides the remarkably high enantioselectivity (Scheme 6.7). Interestingly enough, no reaction occurs if MS 4A is not present during the preparation stage of the chiral catalyst, and MS 4A affects the subsequent allylation reaction. MS 4A dried for 12 h at 250 °C and 0.1 Torr was recommended. Keck reported that addition of $\text{CF}_3\text{CO}_2\text{H}$ or $\text{CF}_3\text{SO}_3\text{H}$ strongly accelerates the reactions catalysed by BINOL–Ti(OPr^{*i*})₂ complex (**2**).⁵⁰



Scheme 6.7



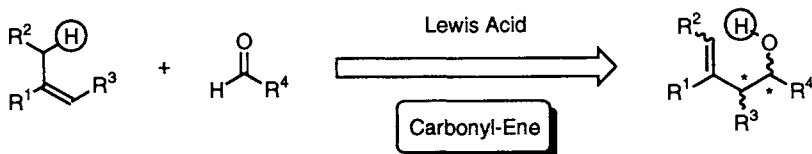
Scheme 6.8

The BINOL–Ti complex-catalysed allylsilane addition to aliphatic and aromatic aldehydes has been reported by Carreira.⁵³ The catalyst is prepared

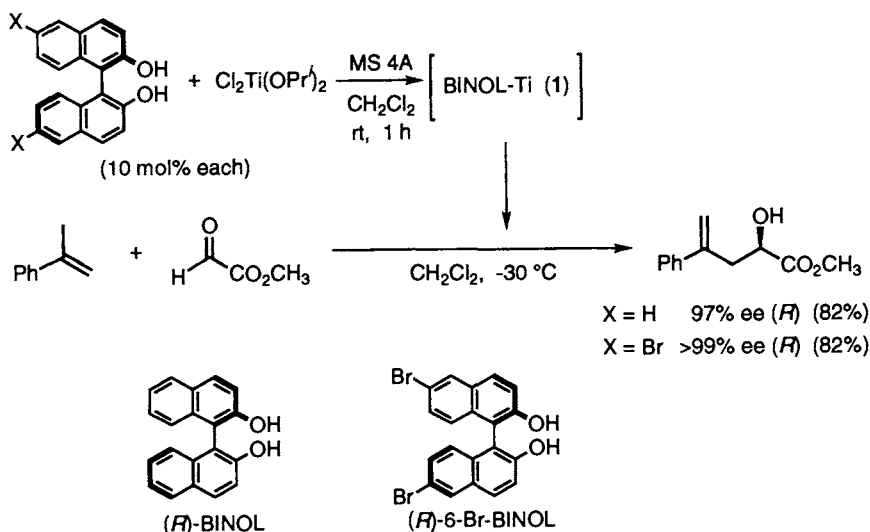
from BINOL and polymeric TiF_4 (Scheme 6.8). The presence of a small amount of CH_3CN is crucial to attain not only high catalytic activity but also high enantioselectivity.

3. Carbonyl-ene reaction

The class of ene reactions which involves a carbonyl compound as the enophile, which we refer to as the 'carbonyl-ene reaction',^{54–57} constitutes a useful synthetic method for the construction of carbon skeleton (Scheme 6.9).



We have been investigating the possibility of stereocontrol in carbonyl-ene reactions promoted by a stoichiometric or catalytic amount of various Lewis acids.^{58–60} In particular, we have developed a chiral titanium catalyst for the glyoxylate-ene reaction which provides the α -hydroxy esters of biological and synthetic importance^{61–64} in an enantioselective fashion (Scheme 6.10).^{65–70} Various chiral titanium catalysts were screened.⁷¹ The best result was obtained with the titanium catalyst (1) prepared *in situ* in the presence of MS

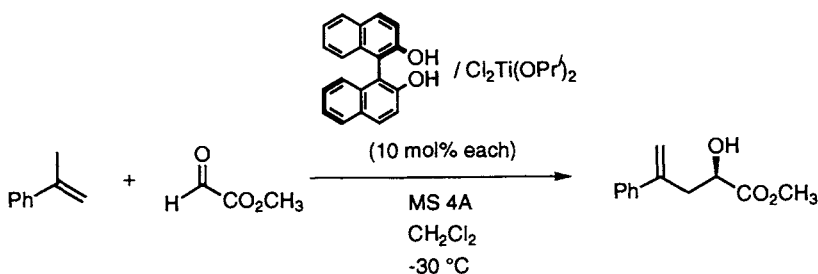


4A from di-isopropoxytitanium dihalides ($X_2Ti(OPr^i)_2$: $X = Br^{24}$ or Cl^{25}) and optically pure BINOL or 6-Br-BINOL^{72–74} (Protocol 3) (These two ligands are now commercially available in (*R*)- and (*S*)-forms.) The remarkable levels of enantioselectivity and rate acceleration observed with these BINOL–Ti catalysts (**1**) stem from the favourable influence of the inherent C_2 symmetry and the higher acidity of BINOLs compared to those of aliphatic diols. The reaction is applicable to a variety of 1,1-disubstituted olefins to provide the ene products in extremely high enantiomeric excess (ee) (Table 6.2). In the reactions with mono- and 1,2-disubstituted olefins, however, no ene product was obtained.

Protocol 3.

Asymmetric carbonyl-ene reaction catalysed by BINOL-derived titanium complex. Catalytic asymmetric synthesis of α -hydroxyester^{24,68,69} (Scheme 6.11)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 6.11

Equipment

- A pre-dried, three-necked, round-bottomed flask (100 mL) equipped with a magnetic stirring bar, a dropping funnel, a three-way stopcock and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- Dry gas-tight syringe
- Thermometers for low temperature

Materials

- Molecular sieves 4A,^a powder, <5 mm, activated, 2.0 g
- (*R*)-(+)- or (*S*)-(–)-1,1'-bi-2-naphthol (FW 286.3), 100 mg, 0.35 mmol
- 0.2 M solution of di-isopropoxytitanium(IV) dibromide in toluene, 1.75 mL, 0.35 mmol
- α -Methylstyrene (FW 118.2), 14.0 mL, 108 mmol
- Freshly distilled methyl glyoxylate^b (FW 88.1), 6.16 g, 70 mmol
- Dry dichloromethane, 45 mL

irritant, hygroscopic

irritant

moisture sensitive, irritant

corrosive, lachrymator

lachrymator, irritant

toxic, irritant

6: Chiral titanium complexes for enantioselective catalysis

1. Add the powder molecular sieves 4A (2.0 g) and (*R*)-(+)-1,1'-bi-2-naphthol (100 mg, 0.35 mmol) to the three-necked flask, purge with argon, and suspend with CH₂Cl₂ (20 mL). Stir for 15 min at room temperature.
2. Add a 0.2 M toluene solution of di-isopropoxytitanium dibromide (0.33 mL, 0.10 mmol) (prepared as described in Protocol 1) by syringe at room temperature.
3. After stirring for 1 h at room temperature, cool to -30°C. Add the α -methylstyrene (14.0 mL, 108 mmol) in CH₂Cl₂ (5 mL) by syringe. Add the methyl glyoxylate (6.16 g, 70 mmol) dissolved in CH₂Cl₂ (20 mL) dropwise over 30 min.
4. Stir for 6 h at -30°C. The reaction temperature must be kept in the range of -35°C to -30°C. The progress of the reaction is monitored by TLC.
5. Pour the CH₂Cl₂ solution into a beaker containing 10 mL of saturated NaHCO₃ and 50 mL of ether. Filter off molecular sieves 4A through a pad of Celite and rinse the filter cake three times with ethyl acetate. Separate the phases and extract the aqueous phase three times with 80 mL of ethyl acetate. Wash the combined organic phases twice with 50 mL of brine, dry over magnesium sulfate, and filter the solution.
6. Concentrate the crude product under vacuum and distil the residue at 0.2 mmHg to collect 12.1 g (84%) of methyl 2-hydroxy-4-phenyl-4-pentenoate boiling at 105–106°C. The product displays the appropriate spectral characteristics and high resolution mass spectral data.
7. The enantiomeric purity is determined by the HPLC analysis using a CHIRALPAK AS as a chiral stationary phase column with 25% *i*-PrOH/hexane as a mobile phase (94% ee).

^a Purchased from Aldrich Chemical Company, Inc. and used without activation.

^b Immediately prior use, the methyl glyoxylate is depolymerized by vacuum distillation from phosphorus pentoxide (10% weight) b.p. 62°C/60 mmHg, bath temperature 120–140°C.

This limitation has been overcome by the use of vinylic sulfides and selenides instead of mono- and 1,2-disubstituted olefins. With these substrates, the ene products are formed with virtually complete enantioselectivity and high diastereoselectivity.⁷⁵ The synthetic utility of the vinylic sulfide and selenide approach is exemplified by the synthesis of enantiopure (*R*)-(-)-ipsdienol, an insect aggregation pheromone (Scheme 6.12).^{76–78}

The synthetic potential of the asymmetric catalytic carbonyl-ene reaction depends greatly on the functionality which is possible in the carbonyl enophile. However, the types of enophile that can be used in the asymmetric catalytic ene reaction have previously been limited to aldehydes such as glyoxylate^{68,69} and chloral.⁷⁰ Thus, it is highly desirable to develop other types of carbonyl enophiles to provide enantio-enriched molecules with a wider range

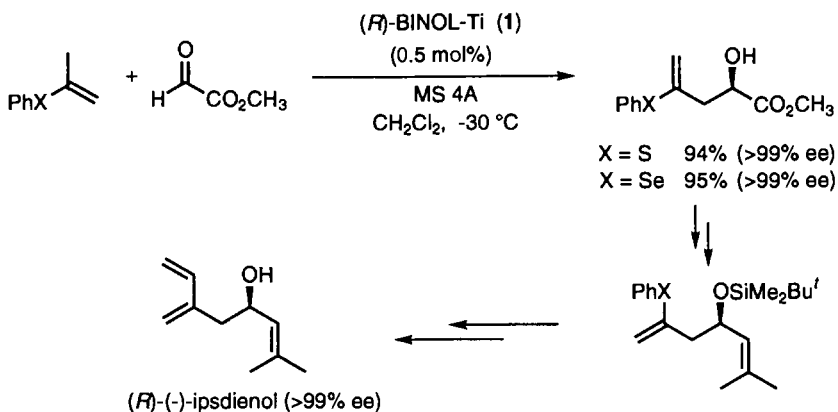
Table 6.2 Asymmetric catalytic glyoxylate-ene reactions with various olefins^a

| Run | olefin | X ₂ Ti(OPr ^{<i>i</i>}) ₂ (X) | catalyst (mol%) | time (h) | products | %yield | %ee ^b |
|-----|--------|---|--------------------|-------------|----------|--------|--------------------------|
| A | | Cl | 10 | 8 | | 72 | 95 <i>R</i> |
| | | Cl | 10 | 8 | | 68 | 95 <i>S</i> ^c |
| | | Cl | 1.0 | 8 | | 78 | 93 <i>R</i> |
| | | Br | 10 | 3 | | 87 | 94 <i>R</i> |
| B | | Cl | 1.0 | 8 | | 97 | 97 (<i>R</i>) |
| | | Br | 1.0 | 3 | | 98 | 95 (<i>R</i>) |
| C | | Cl | 10 | 8 | | 82 | 97 (<i>R</i>) |
| | | Br | 5 | 3 | | 89 | 98 (<i>R</i>) |
| D | | Cl | 10 | 8 | | 87 | 88 (<i>R</i>) |
| | | Br | 5 | 3 | | 92 | 89 (<i>R</i>) |

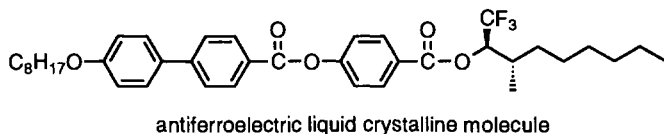
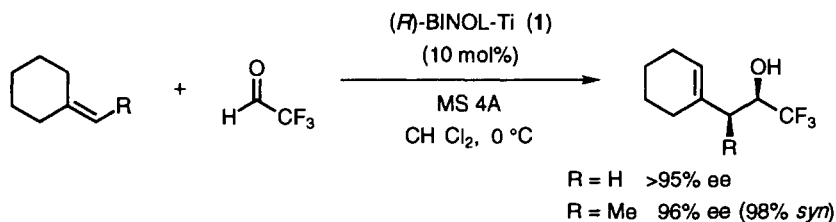
^a All reactions were carried out using 1.0 mmol of methyl glyoxylate, 2.0 mmol of olefin, and indicated amount of BINOL-Ti complex at -30 °C.

^b Determined by LIS-NMR analysis after conversion to the corresponding α-methoxy esters. The configuration in parentheses could be assigned by the similarity in shift pattern seen in the LIS-NMR spectra using (+)-Eu(dppm)₃ as a chiral shift reagent.

^c (S)-BINOL was used instead of the (*R*)-counterpart.

**Scheme 6.12**

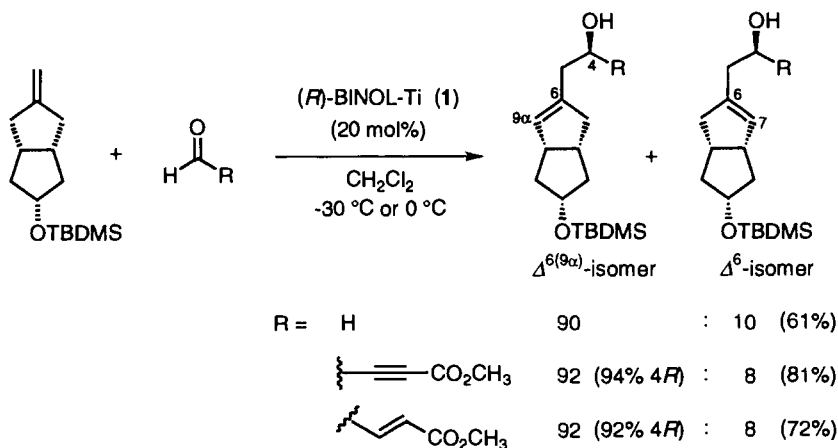
of functionalities. We have developed an asymmetric catalytic fluoral-ene reaction,^{79–81} which provides an efficient approach for the asymmetric synthesis of some fluorine-containing compounds of biological and synthetic importance.⁸² The reaction of fluoral with 1,1-disubstituted and trisubstituted olefins proceeds quite smoothly under the catalysis by the BINOL-Ti complex (**1**) to provide the corresponding homoallylic alcohol with extremely high enantioselectivity (>95% ee) and *syn*-diastereoselectivity (>90%) (Scheme 6.13).



Scheme 6.13

The sense of asymmetric induction in the fluoral-ene reaction is exactly the same as observed for the glyoxylate-ene reaction; (R) -BINOL-Ti (**1**) provides the (R) - α - CF_3 alcohol. The *syn*-diastereomers of α -trifluoromethyl- β -methyl-substituted compounds thus synthesized with *two stereogenic centres* show anti-ferroelectric properties preferentially to the *anti*-diastereomers.^{83–86}

The BINOL-Ti catalysis can also be used for the carbonyl-ene reaction with formaldehyde or vinyl and alkynyl analogues of glyoxylates in an asymmetric catalytic desymmetrization (see below) approach to the asymmetric synthesis of isocarbacycline analogues (Scheme 6.14).^{87,88}

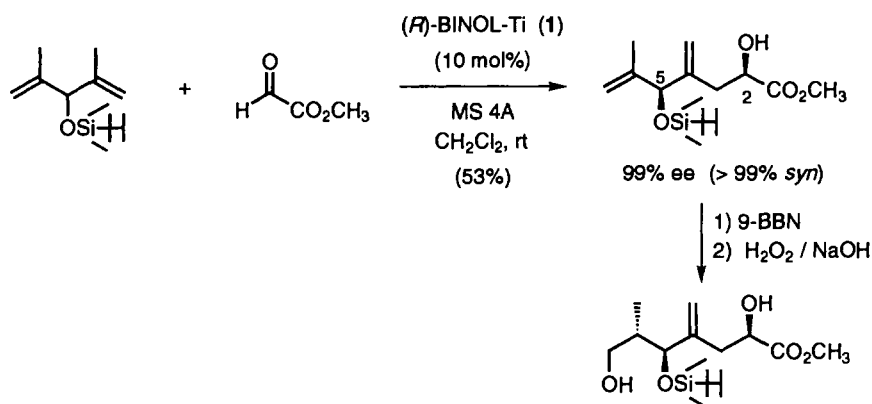


Scheme 6.14

4. Asymmetric catalytic desymmetrization

Desymmetrization of an achiral, symmetrical molecule through a catalytic process is a potentially powerful but relatively unexplored concept for

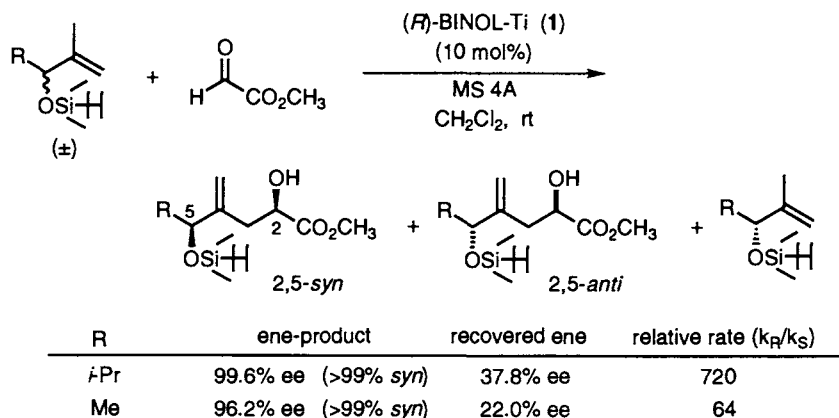
asymmetric synthesis. Although the ability of enzymes to differentiate between enantiotopic functional groups is well known,⁸⁹ little has been explored on a similar ability of non-enzymatic catalysts, particularly for carbon–carbon bond-forming processes. The desymmetrization by the catalytic glyoxylate–ene reaction of prochiral ene substrates with the planar symmetry provides an efficient access to remote⁹⁰ and internal⁹¹ asymmetric induction which is otherwise difficult to attain (Scheme 6.15).⁹² The (2*R*,5*S*)-*syn*-product is obtained in >99% ee along with more than 99% diastereoselectivity. The diene thus obtained can be transformed to a more functionalized compound in a regioselective and diastereoselective manner.



Scheme 6.15

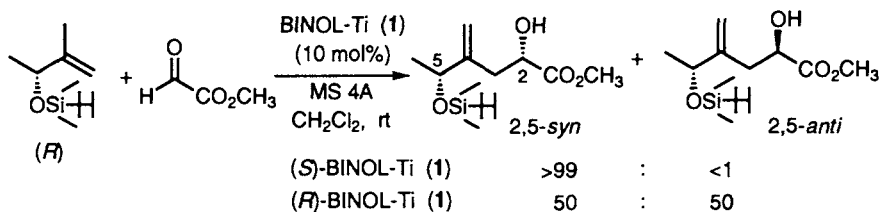
5. Kinetic optical resolution

On the basis of desymmetrization concept, the kinetic optical resolution of a racemic substrate^{93,94} might be recognized as an intermolecular version of desymmetrization. The kinetic resolution of a racemic allylic ether by the glyoxylate–ene reaction also provides an efficient access to remote but relative⁹¹ asymmetric induction. The reaction of allylic ethers catalysed by the (*R*)-BINOL-derived complex (1) provides the 2*R*,5*S*-*syn*-products with >99% diastereoselectivity along with more than 95% ee (Scheme 6.16). The high diastereoselectivity, coupled with the high ee, strongly suggests that the catalyst/glyoxylate complex efficiently discriminates between the two enantiomeric substrates to accomplish the effective kinetic resolution. In fact, the relative rates between the reactions of the either enantiomers, calculated by the equation $\{1n[(1-c)(1-ee_{\text{recov}})] \times [1n[(1-c)(1+ee_{\text{recov}})]]^{-1}$, $c = (ee_{\text{recov}}) \times (ee_{\text{recov}} + ee_{\text{prod}})^{-1}$, $0 < c$, $ee < 1$ where c is the fraction of consumption), were approx. 700 for $R = i\text{-Pr}$ and 65 for $R = \text{Me}$. As expected, the double asymmetric induction^{95,96} in the reaction of (*R*)-ene component using the



Scheme 6.16

catalyst (*S*)-1 ('matched' catalytic system) leads to the complete (>99%) 2,5-syn-diastereoselectivity in high chemical yield, whereas the reaction of (*R*)-ene using (*R*)-1 ('mis-matched' catalytic system) produces a diastereomeric mixture in quite low yield (Scheme 6.17).

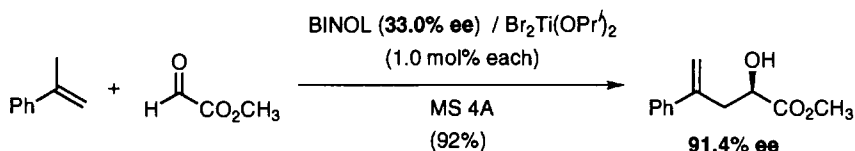


Scheme 6.17

6. Positive non-linear effect of non-racemic catalysts

Deviation from the linear relationship, namely 'non-linear effect' (NLE) is sometimes observed between the enantiomeric purity of chiral catalysts and the optical yields of the products. Among, the convex deviation, which Kagan^{97,98} and Mikami^{99,100} independently refer to as positive non-linear effect (abbreviated as (+)-NLE (asymmetric amplification¹⁰¹)) has attracted current attention to achieve higher level of asymmetric induction than the enantio-purity of the non-racemic (partially resolved) catalysts.^{97,98,101-103} In turn, (-)-NLE stands for the opposite phenomenon of concave deviation, namely negative non-linear effect.

We have observed a remarkable level of (+)-NLE in the catalytic ene reaction. For instance, in the glyoxylate-ene reaction, the use of a catalyst prepared from BINOL of 33.0% ee provides the ene product with 91.4% ee in 92% chemical yield (Scheme 6.18).^{99,100} The ee thus obtained is not only much



Scheme 6.18

higher than the ee of the BINOL employed, but also very close to the value (94.6% ee) obtained using enantiomerically pure BINOL (Fig. 6.1).

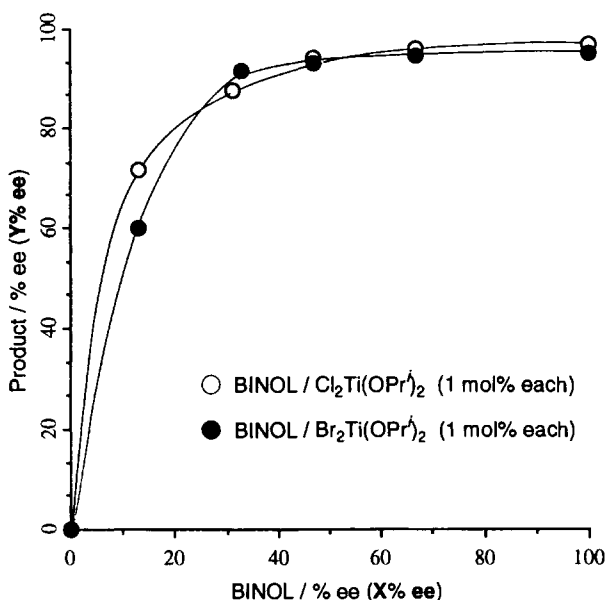


Fig. 6.1. (+)-NLE in asymmetric glyoxylate-ene reaction catalysed by BINOL-Ti complex (2).

7. Enantiomer-selective activation of racemic catalysts

Whereas non-racemic catalysts can generate non-racemic products with or without NLE, racemic catalysts inherently produce only racemic products. A strategy whereby a racemic catalyst is enantiomer-selectively de-activated by a chiral molecule has been shown to yield non-racemic products.^{104,105} However, the level of asymmetric induction does not exceed the level attained by the enantiopure catalyst (Fig. 6.2a). Recently, 'chiral poisoning'¹⁰⁶ has been named for such a *de-activating* strategy. In contrast, we have reported an alternative but conceptually opposite strategy to asymmetric catalysis by

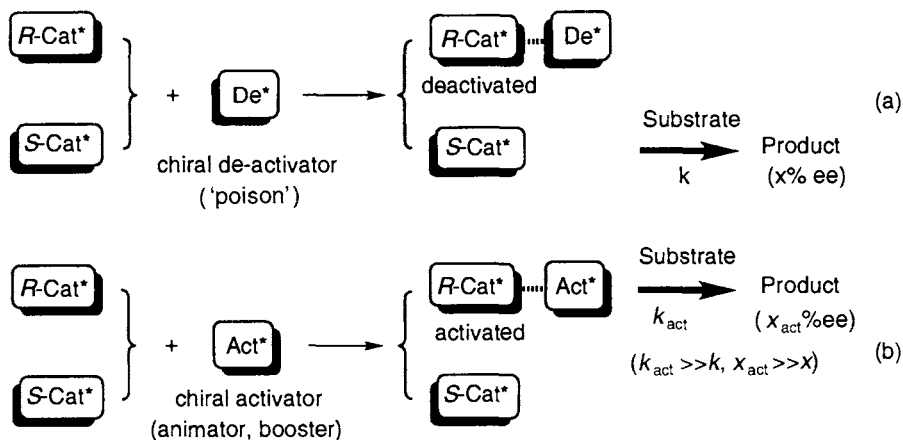
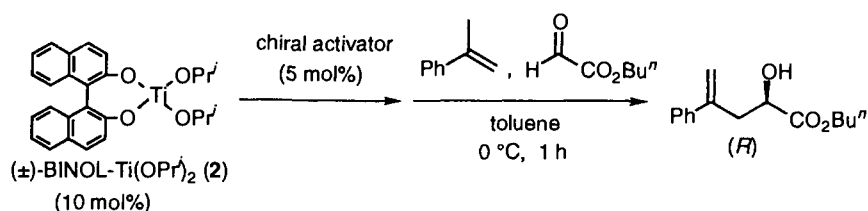


Fig. 6.2. Asymmetric de-activation (a) and asymmetric activation (b) of racemic catalysts.

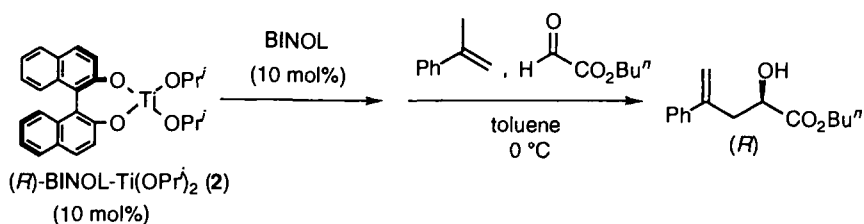
racemic catalysts. A *chiral activator* selectively activates one enantiomer of a racemic chiral catalyst. Higher enantioselectivity might be attained than that achieved by an enantio-pure catalyst ($\% \text{ ee}_{\text{act}} \gg \% \text{ ee}_{\text{enantio-pure}}$), in addition to a higher level of catalytic efficiency ($k_{\text{act}} \gg k_{\text{enantio-pure}}$) (Fig. 6.2b). However, this still remains to be examined.

Catalysis with racemic $\text{BINOL-Ti(OPr}^i)_2$ (**2**) achieves extremely high enantioselectivity by adding another diol for the enantiomer-selective activation (Table 6.3).^{107–110} Significantly, a remarkably high enantioselectivity (90% ee, *R*) was achieved using just a half-molar amount (5 mol%) of (*R*)-BINOL activator added to a *racemic* (\pm)-BINOL- $\text{Ti(OPr}^i)_2$ complex (**2**) (10 mol%).

The activation of the enantiopure (*R*)-BINOL- $\text{Ti(OPr}^i)_2$ catalyst (**2**) was investigated by further addition of (*R*)-BINOL (Table 6.4). The reaction proceeded quite smoothly to provide the carbonyl-ene product in higher chemical yield (82%) and enantioselectivity (97% ee) than without additional BINOL (95% ee, 20%). Comparing the results of enantiomer-selective activation of the racemic catalyst (90% ee, *R*) with those of the enantiopure catalyst (with (97% ee, *R*) or without activator (95% ee, *S*)), the reaction catalysed by the (*R*)-BINOL- $\text{Ti(OPr}^i)_2$ /(*R*)-BINOL complex (**2'**) is calculated to be 26.3 times as fast as that catalysed by the (*S*)-BINOL- $\text{Ti(OPr}^i)_2$ (**2**) in the racemic case (Fig. 6.3). Indeed, kinetic studies according to a rapid-quench GC analysis show that the reaction catalysed by the (*R*)-BINOL- $\text{Ti(OPr}^i)_2$ /(*R*)-BINOL complex (**2'**) is 25.6 times as fast as that catalysed by the (*R*)-BINOL- $\text{Ti(OPr}^i)_2$ (**2**). These results imply that the racemic (\pm)-BINOL- $\text{Ti(OPr}^i)_2$ (**2**) and half-molar amount of (*R*)-BINOL assemble preferentially into the (*R*)-BINOL- $\text{Ti(OPr}^i)_2$ /(*R*)-BINOL complex (**2'**) and unchanged (*S*)-BINOL- $\text{Ti(OPr}^i)_2$ (**2**). In contrast, the enantiomeric

Table 6.3 Enantiomer selective activation of racemic BINOL-Ti(OPr^{*i*})₂ (2)

| Run | Chiral activator | %yield | % ee |
|----------------|------------------|--------|------|
| 1 | none | 1.6 | 0 |
| 2 | | 20 | 0 |
| 3 | | 38 | 81 |
| 4 | | 52 | 90 |
| 5 ^a | | 35 | 80 |

^a 2.5 mol% of (*R*)-BINOL was used as a chiral activator.**Table 6.4** Asymmetric activation of enantiopure (*R*)-BINOL-Ti(OPr^{*i*})₂ (2)

| Run | BINOL | Time (min) | %yield | % ee |
|-----|--------------------|------------|--------|------|
| 1 | none | 60 | 20 | 95 |
| 2 | | 1 | 1.6 | 95 |
| 3 | (<i>R</i>)-BINOL | 60 | 82 | 97 |
| 4 | | 1 | 41 | 97 |
| 5 | | 0.5 | 24 | 97 |
| 6 | (<i>S</i>)-BINOL | 60 | 48 | 86 |
| 7 | | 0.5 | 2.6 | 86 |
| 8 | (±)-BINOL | 60 | 69 | 96 |

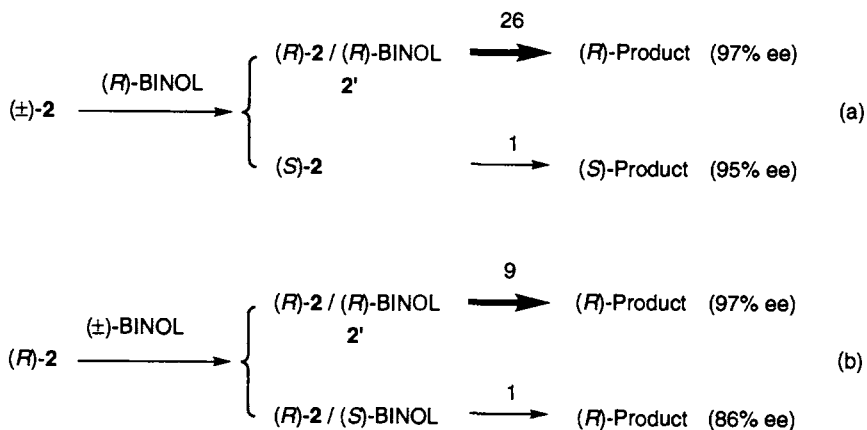


Fig. 6.3. Kinetic features of asymmetric activation of BINOL-Ti(OPrⁱ)₂ (3).

form of the additional chiral ligand ((*S*)-BINOL) activates the (*R*)-BINOL-Ti(OPrⁱ)₂ (2) to a smaller degree, thus providing the carbonyl-ene product in lower optical (86% ee, *R*) and chemical (48%) yields than (*R*)-BINOL does.

The great advantage of asymmetric activation of the racemic BINOL-Ti(OPrⁱ)₂ complex (2) is highlighted in a catalytic version (Table 6.3, Run 5). High enantioselectivity (80.0% ee) is obtained by adding less than the stoichiometric amount (0.25 molar amount) of additional (*R*)-BINOL. A similar phenomenon on enantiomer-selective activation has been observed in aldol and (hetero) Diels-Alder reactions catalyzed by a racemic BINOL-Ti(OPrⁱ)₂ catalyst (2) (see below).

Another possibility was explored using racemic BINOL as an activator. Racemic BINOL was added to the (*R*)-BINOL-Ti(OPrⁱ)₂ (2), giving higher yield and enantioselectivity (96% ee, 69%) than those obtained by the original catalyst (*R*)-BINOL-Ti(OPrⁱ)₂ (2) without additional BINOL (95% ee, 20%). Comparing the results (96% ee, *R*) by the racemic activator with those of enantiopure catalyst, (*R*)-BINOL-Ti(OPrⁱ)₂/*R*-BINOL (2') (97% ee, *R*) or (*R*)-BINOL-Ti(OPrⁱ)₂/*S*-BINOL (86% ee, *R*), the reaction catalysed by the (*R*)-BINOL-Ti(OPrⁱ)₂ catalyst/*R*-BINOL complex (2') is calculated to be 8.8 times as fast as that catalysed by the (*R*)-BINOL-Ti(OPrⁱ)₂/*S*-BINOL (Fig. 6.3). A rapid-quench GC analysis revealed the reaction catalysed by the (*R*)-BINOL-Ti(OPrⁱ)₂/*R*-BINOL complex (2') to be 9.2 times as fast as that catalysed by the (*R*)-BINOL-Ti(OPrⁱ)₂/*S*-BINOL.

8. Ene cyclization

Conceptually, intramolecular ene reactions¹¹¹⁻¹¹⁴ (ene cyclizations) can be classified into six different modes (Fig. 6.4).^{54,115} In the ene cyclizations, the

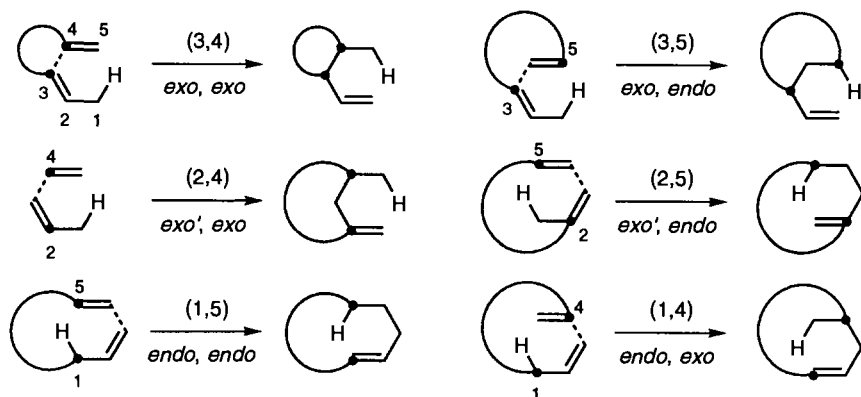
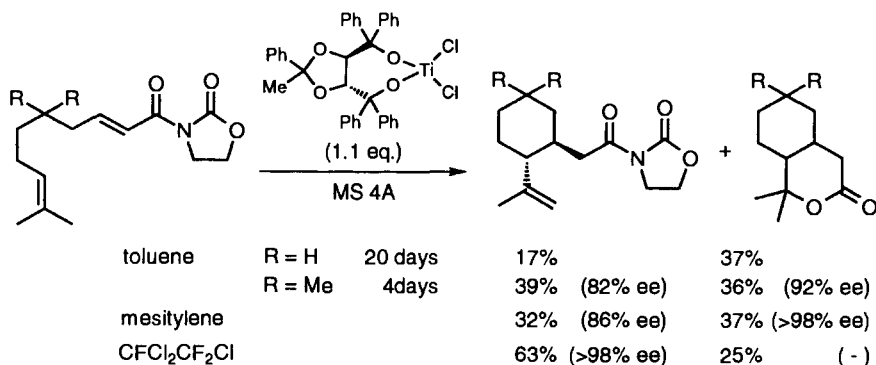


Fig. 6.4. Classification of ene cyclizations.

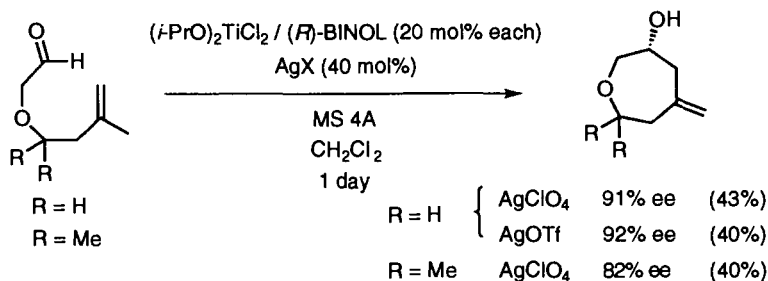
carbon numbers where the tether connects the [1,5]-hydrogen shift system, are expressed in (m,n) type. A numerical prefix stands for the forming ring size.

Asymmetric catalysis of ene reactions was initially explored in the intramolecular cases, since the intramolecular versions are much more facile than their intermolecular counterparts. The first example of an enantioselective 6-(3,4) carbonyl-ene cyclization was reported using a BINOL-derived zinc reagent.^{116,117} However, this was successful only when using an excess of the zinc reagent (at least 3 equivalents). Recently, an enantioselective 6-(3,4) olefin-ene cyclization has been developed using a stoichiometric amount of a TADDOL-derived chiral titanium complex (Scheme 6.19).¹¹⁸ In this ene reaction, a hetero Diels–Alder product was also obtained, the ratio depending critically on the solvent system employed. In both cases, geminal disubstitution is required in order to obtain high ee's values. However, neither reaction constitutes an example of a truly catalytic asymmetric ene cyclization.

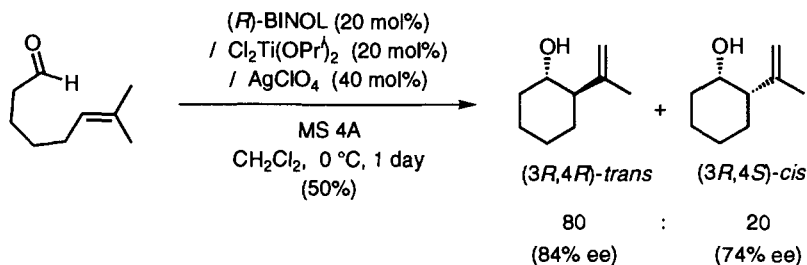


Scheme 6.19

We reported the first examples of asymmetric catalysis of intramolecular carbonyl-ene reactions of types (3,4) and (2,4), using the BINOL-derived titanium complex (**1**).^{115,119} The catalytic 7-(2,4) carbonyl-ene cyclization gives the oxepane with high ee, and gem-dimethyl groups are not required (Scheme 6.20). In a similar catalytic 6-(3,4) ene cyclization, the *trans*-tetrahydropyran is preferentially obtained with high ee (Scheme 6.21). The sense of asymmetric induction is exactly the same as observed for the glyoxylate-ene reaction: the (*R*)-BINOL-Ti catalyst provides the (*R*)-cyclic alcohol. Therefore, the chiral BINOL-Ti catalyst works efficiently for both the chiral recognition of the enantioface of the aldehyde and the discrimination of the diastereotopic protons of the ene component in a truly catalytic fashion.



Scheme 6.20



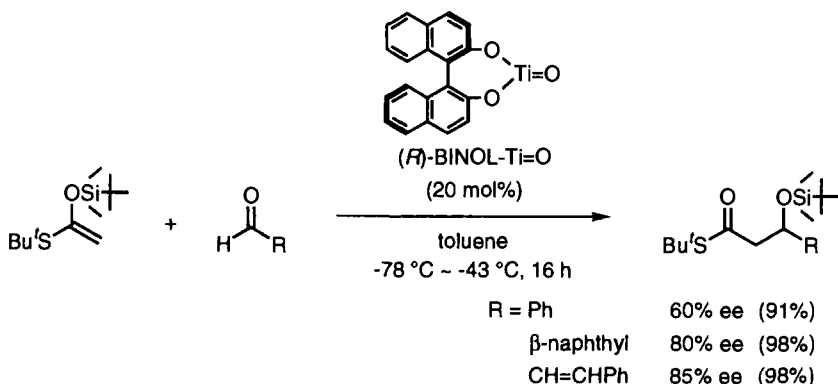
Scheme 6.21

9. Aldol reaction

The aldol reaction constitutes one of the most fundamental bond construction processes in organic synthesis.^{120,121} Therefore, much attention has been focused on the development of asymmetric catalysts for aldol reactions using silyl enol ethers of ketones or esters as storable enolate components (the Mukaiyama aldol condensation).

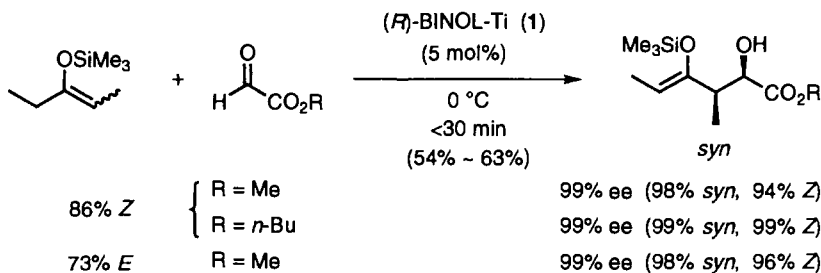
Reetz reported the catalysis by BINOL-TiCl₂ of aldol reactions with aliphatic aldehydes.¹²² In his case, BINOL-TiCl₂ was prepared by treatment of the lithium salt of BINOL with TiCl₄ in ether. After removal of the ether, the

residue was treated with dry benzene and the solid was separated under nitrogen. Removal of the solvent provided the red-brown complex, which was used as the catalyst for the aldol reaction to give 8% ee. Later, Mukaiyama reported the use of BINOL-Ti oxide prepared from $(i\text{-PrO})_2\text{Ti=O}$ and BINOL, giving moderate to high levels of enantioselectivity (Scheme 6.22).^{123,124}



Scheme 6.22

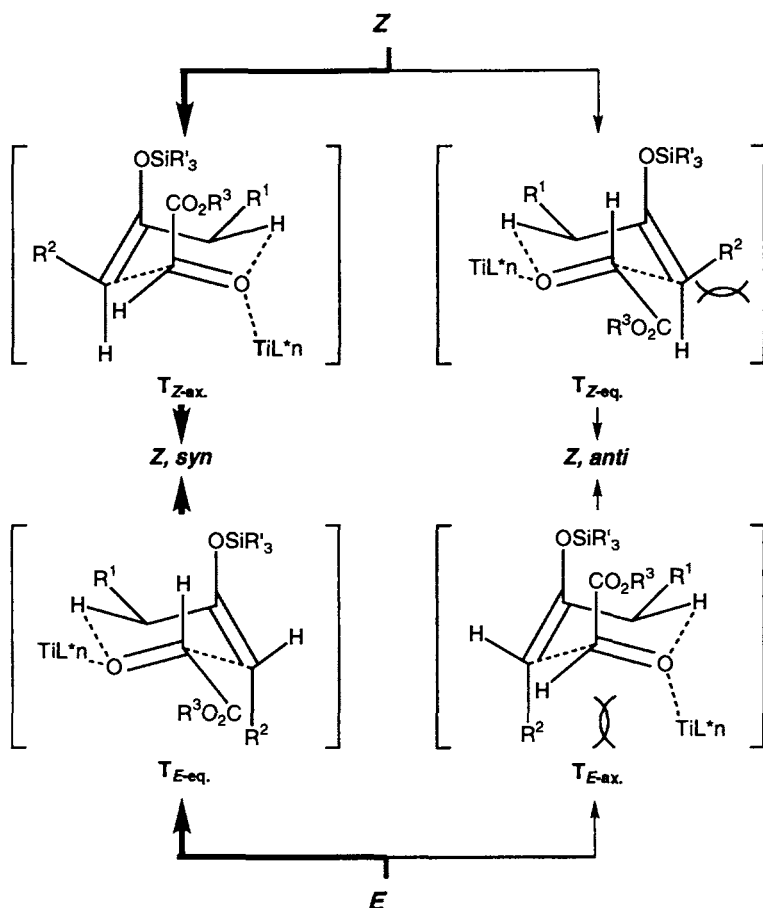
We have found that the BINOL-derived titanium complex serves as an efficient catalyst for the Mukaiyama-type aldol reaction of ketone silyl enol ethers with good control of both absolute and relative stereochemistry (Scheme 6.23).¹²⁵ Surprisingly, however, the aldol products were obtained in the silyl enol ether (ene product) form, with high *syn*-diastereoselectivity from either geometrical isomer of the starting silyl enol ethers.



Scheme 6.23

It appears likely that the reaction proceeds through an ene reaction pathway. Such an ene reaction pathway has not been previously recognized as a possible mechanism in the Mukaiyama aldol condensation. Usually an acyclic antiperiplanar transition state model has been used to explain the formation of the *syn*-diastereomer from either (*E*)- or (*Z*)-silyl enol ethers.^{126,127} How-

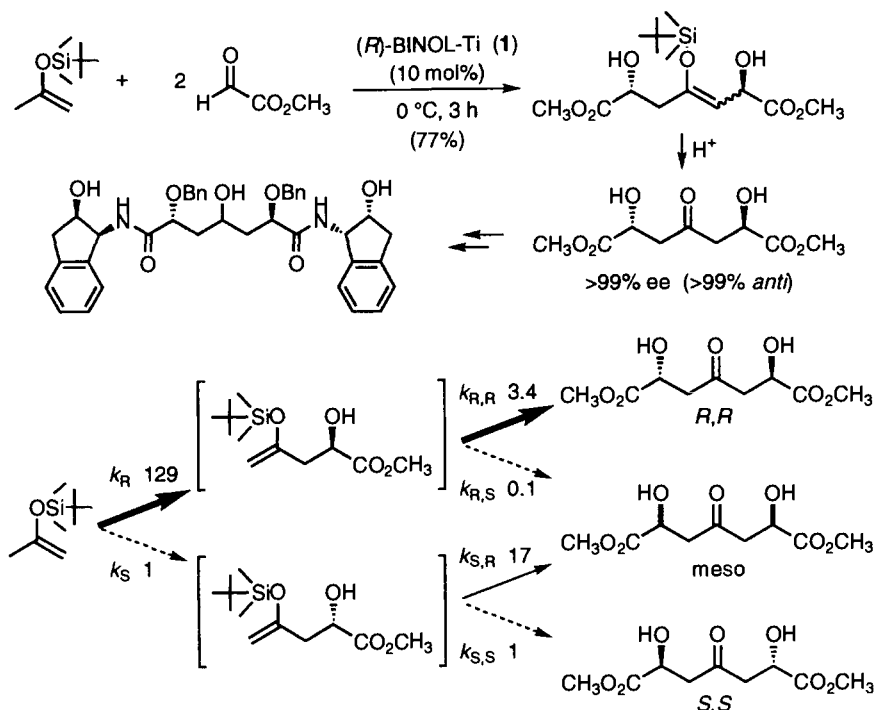
ever, the cyclic ene mechanism now provides another rationale for the *syn*-diastereoselection irrespective of the enol silyl ether geometry (Scheme 6.24).



Scheme 6.24

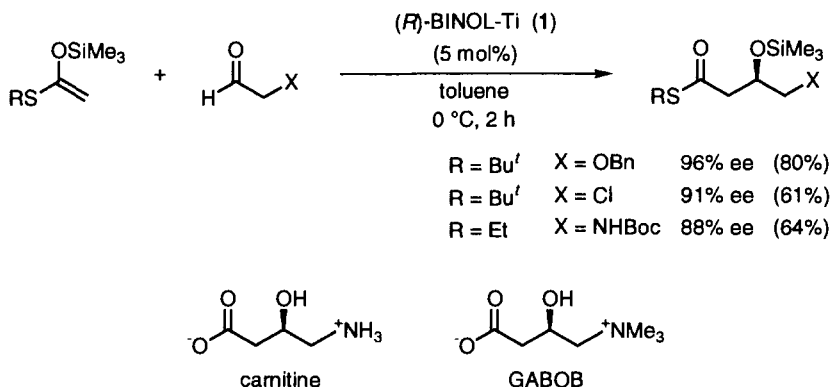
The aldol reaction of a silyl enol ether proceeds in double and two-directional fashion, upon addition of an excess amount of an aldehyde, to give the silyl enol ether in 77% isolated yield and in more than 99% ee and 99% de (Scheme 6.25).¹²⁸ The present asymmetric catalytic aldol reaction is characterized by a kinetic amplification phenomenon of the product chirality on going from the one-directional aldol intermediate to the two-directional product. Further transformation of the *pseudo* C₂ symmetric product whilst still being protected as the silyl enol ether leads to a potent analogue of an HIV protease inhibitor.

The silatropic ene pathway, that is direct silyl transfer from an enol silyl ether to an aldehyde, may be involved as a possible mechanism in the



Scheme 6.25

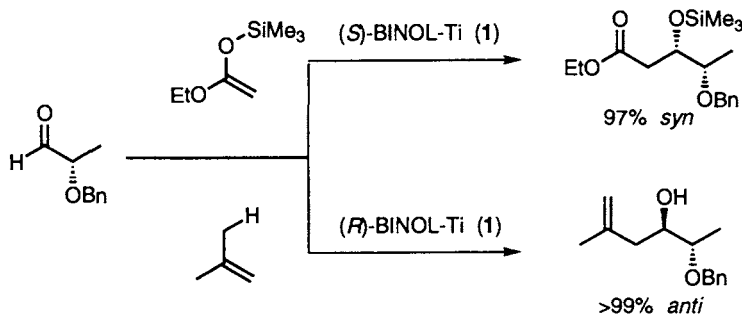
Mukaiyama aldol-type reaction. Indeed, *ab initio* calculations show the silatropic ene pathway involving the cyclic (boat and chair) transition states for the BH_3 -promoted aldol reaction of the trihydrosilyl enol ether derived from acetaldehyde with formaldehyde to be favoured.¹²⁹ Recently, we have reported the possible intervention of a silatropic ene pathway in the asymmetric



Scheme 6.26

catalytic aldol-type reaction of silyl enol ethers of thioesters.¹³⁰ Chloro and amino compounds thus obtained are useful intermediates for the synthesis of carnitine and GABOB (Scheme 6.26).^{131,132}

There is a dichotomy in the sense of *syn*- vs. *anti*-diastereofacial preference, dictated by the bulkiness of the migrating group.¹²⁹ The sterically demanding silyl group shows *syn*-diastereofacial preference but the less demanding proton leads to *anti*-preference (Scheme 6.27). The *anti*-diastereoselectivity in carbonyl-ene reactions can be explained by the Felkin–Anh-like cyclic transition state model (**T₁**) (Fig. 6.5). In the aldol reaction, by contrast, the inside-crowded transition state (**T₁'**) is less favourable than **T₂'**, because of the steric repulsion between the trimethylsilyl group and the inside methyl group of aldehyde (**T₁'**). Therefore, the *syn*-diastereofacial selectivity is visualized by the anti-Felkin–Anh-like cyclic transition state model (**T₂'**).



Scheme 6.27

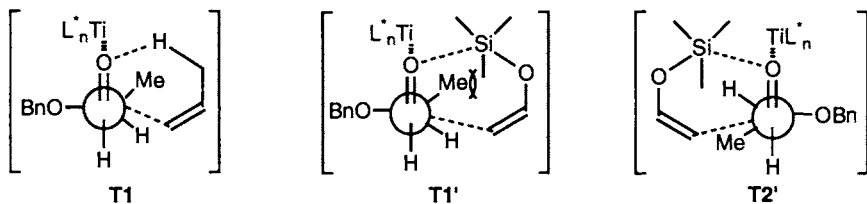
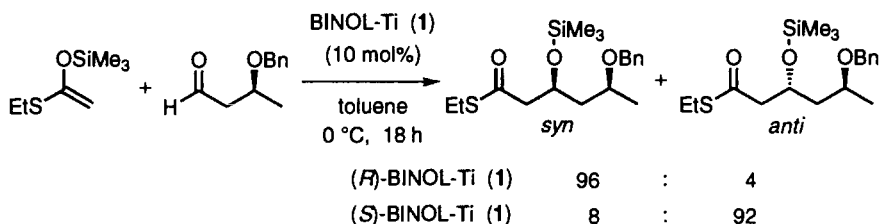


Fig. 6.5. Cyclic transition state models of ene and silatropic ene pathway.

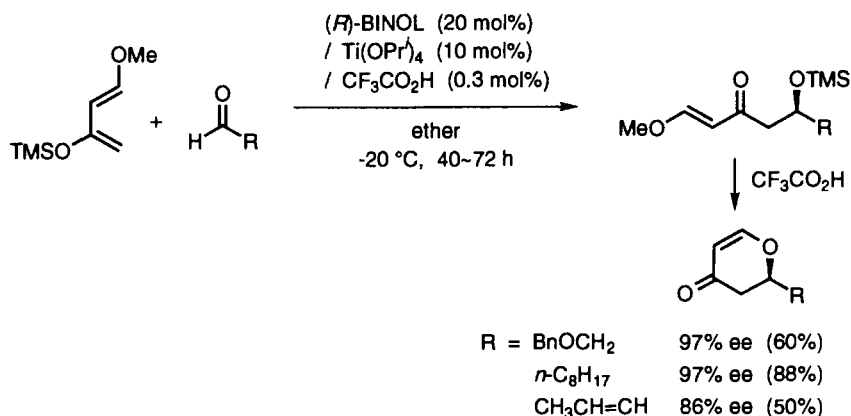
An aldol reaction with chiral β -benzyloxy aldehyde provides a method for the stereodivergent synthesis of both *syn*- and *anti*-diastereomers^{133–136} with high levels of diastereoselectivity dictated primarily by the chirality of a BINOL-Ti catalyst (**1**) rather than a β -benzyloxy aldehyde (Scheme 6.28).^{137–139} The aldol products can be used as useful key intermediates for β -lactone synthesis.¹⁴⁰

Keck¹⁴¹ and Carreira^{142,143} have independently reported catalytic asymmetric Mukaiyama aldol reactions. Keck *et al.* also reported an aldol reaction of α -benzyloxy aldehyde with Danishefsky's diene. The aldol product was trans-



Scheme 6.28

formed to the hetero Diels–Alder type product through acid-catalysed cyclization. In their method, the catalyst is prepared using 1:1 and 2:1 stoichiometry of BINOL and $\text{Ti}(\text{OPr}^i)_4$ (Scheme 6.29).¹⁴⁴ In their cases, oven dried MS 4A is used to generate the catalyst, which they reported to be of BINOL- $\text{Ti}(\text{OPr}^i)_2$ structure, under refluxing conditions.

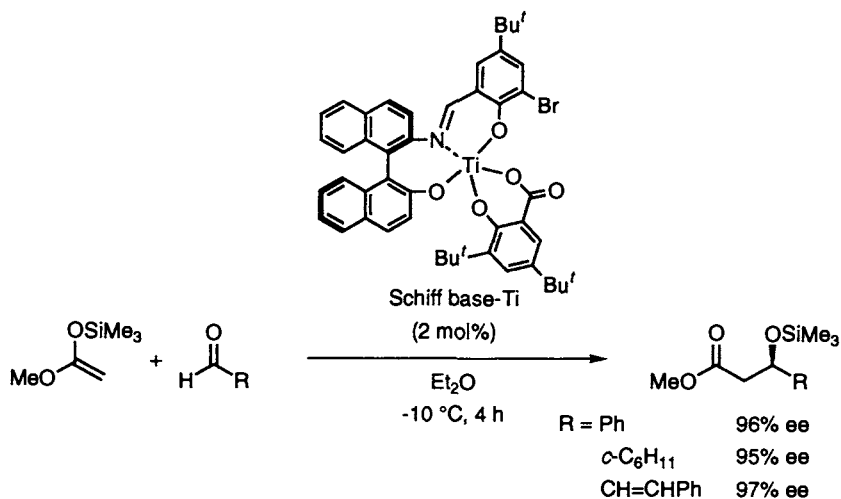


Scheme 6.29

Carreira employed a chiral BINOL-derived Schiff base–titanium complex as a catalyst for aldol reactions with acetate-derived ketene silyl acetals (Scheme 6.30).^{142,143} The catalyst was prepared in toluene in the presence of salicylic acid, which was reported to be crucial to attain a high enantioselectivity.

10. (Hetero) Diels–Alder reaction

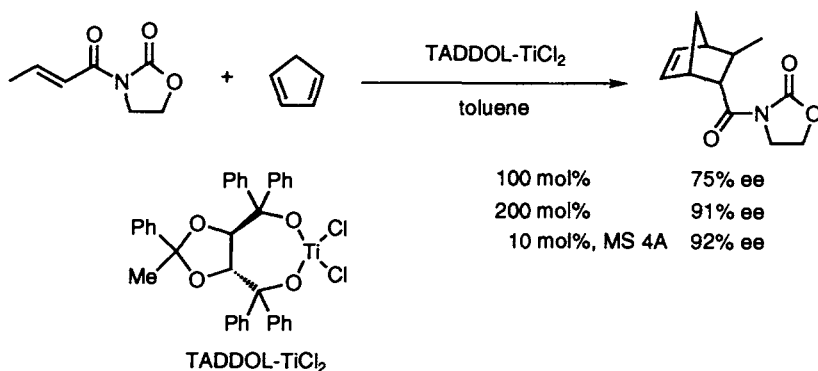
The (hetero) Diels–Alder (D–A) reaction also constitutes one of the most efficient carbon–carbon bond forming processes in the construction of six-membered rings by virtue of its high level of regioselectivity and the high



Scheme 6.30

potential for control of the absolute stereochemistry of the up to four newly created chiral centres.^{145,146}

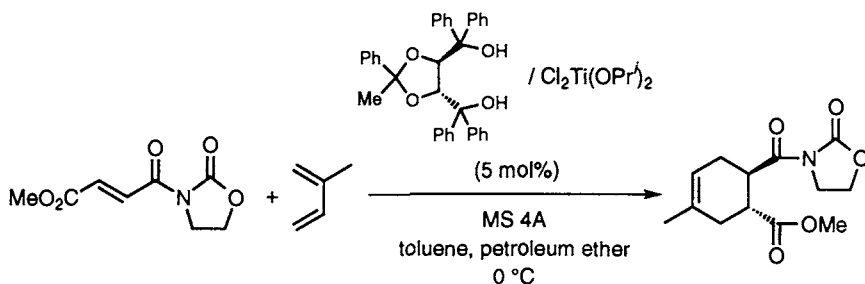
Narasaka has reported that a TADDOL–Ti dichloride—in the presence of MS 4A, acts as a good catalyst for the asymmetric catalytic D–A reactions with oxazolidinone derivatives of acrylates, giving extremely high enantioselectivities (Scheme 6.31) (Protocol 4).^{147–153}



Scheme 6.31

Protocol 4.**Asymmetric Diels–Alder reaction catalysed by TADDOL-derived titanium complex^{147,148} (Scheme 6.32)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 6.32****Equipment**

- A pre-dried, two-necked, round-bottomed flask (30 mL) equipped with a magnetic stirring bar, a three-way stopcock, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- Dry gas-tight syringe
- A pre-dried, two-necked, round-bottomed flask (500 mL) equipped with a three-way stopcock, a magnetic stirring bar, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- Double-ended cannula

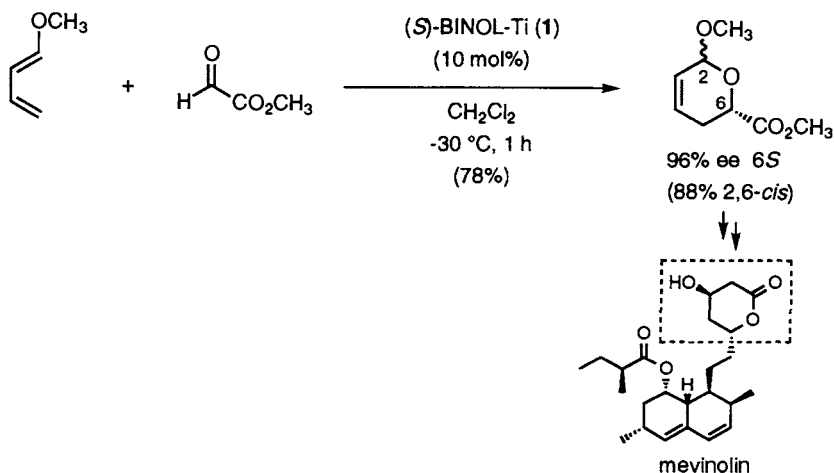
Materials

- Molecular sieves 4A, powder, <5 μm , activated, 3.74 g irritant, hygroscopic
- (2*R*,3*R*)-2,3-*O*-(1-phenylethylidene)-1,1,4,4-tetraphenylbutane-1,2,3,4-tetraol (FW 528.6), 1.32 g, 2.5 mmol
- Di-isopropoxytitanium(IV) dichloride (FW 237.0), 540 mg, 2.3 mmol moisture sensitive, irritant
- 3-((*E*)-3-(Methoxycarbonyl)propenyl)-1,3-oxazolidin-2-one (FW 199.0), 9.10 g, 46 mmol irritant
- Isoprene (FW 68.1), 50 mL cancer suspect agent, flammable liquid
- Dry toluene (distilled from P_2O_5 and from CaH_2 and stored over MS 4A) flammable liquid, toxic
- Dry petroleum ether, 150 mL flammable liquid, toxic

1. Add the di-isopropoxytitanium dichloride (540 mg, 2.3 mmol) (prepared in a similar manner as described in protocol 1) to the two-necked flask (30 mL) under argon flow, then dissolve in toluene (5 mL).
2. Add (2*R*,3*R*)-2,3-*O*-(1-phenylethylidene)-1,1,4,4-tetraphenylbutane-1,2,3,4-tetraol (1.32 g, 2.5 mmol) in toluene (5 mL) by syringe, then stir for 1 h at room temperature.
3. In the separate flask (500 mL), add toluene (175 mL) to the powdered molecular sieves 4A (3.74 g) to give a suspension.

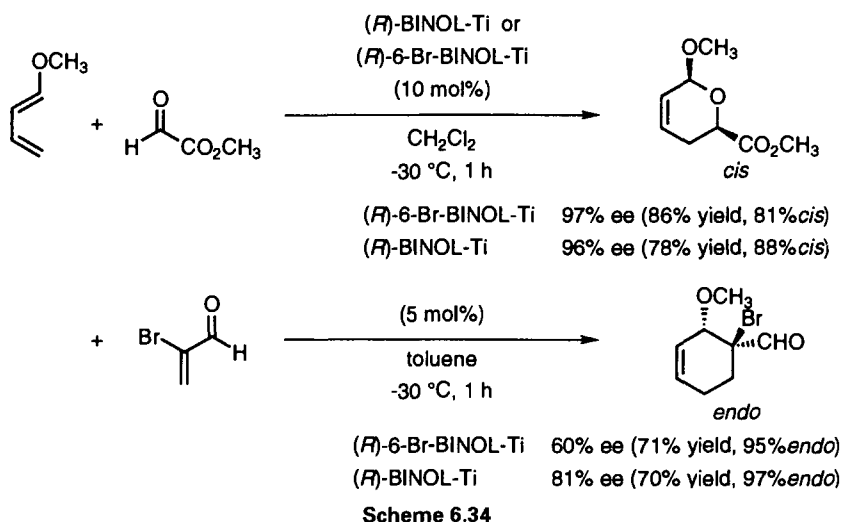
4. Add, *via* the double-ended cannula, the titanium catalyst solution to the suspension of molecular sieves 4A at room temperature.
5. Cool to 0 °C, then add 3-((*E*)-3-(methoxycarbonyl)propenoyl)-1,3-oxazolidin-2-one (9.10 g, 46 mmol), petroleum ether (150 mL), and isoprene (50 mL) in this order.
6. After stirring overnight at 0 °C, add pH 7 phosphate buffer. Filter off molecular sieves 4A through a pad of celite and rinse the filter cake with ethyl acetate. Separate the phases and extract the aqueous phase three times with ethyl acetate. Dry over Na₂SO₄, and filter the solution.
7. Concentrate the crude product under vacuum and purify the crude product by silica-gel column chromatography with hexane-ethyl acetate (4:1) as eluent to obtain 3-(((4'*R*,5'*R*)-1'-methyl-5'-(methoxycarbonyl)cyclohexen-4'-yl)-carbonyl)-1,3-oxazolidin-2-one (11.6 g, 94% yield, 93% ee). The product displays the appropriate spectral characteristics and elemental analysis. The enantiomeric purity is determined by NMR analysis with a chiral NMR shift reagent, Eu(hfc)₃ (MeO signal separated).
8. Recrystallize from hexane-ethyl acetate to give the optically pure product in 64% yield.

We have also reported that the hetero D–A reactions of glyoxylates with 1-methoxy-1,3-butadienes proceed smoothly under catalysis by BINOL–Ti complex to give the *cis*-product with high ee (Scheme 6.33).¹⁵⁴ The hetero D–A products thus obtained can be transformed into monosaccharides.^{155,156} Furthermore, the hetero D–A product can readily be converted into the lactone portion of HMG-Co A inhibitors such as mevinolin or compactin¹⁵⁷ in few steps.

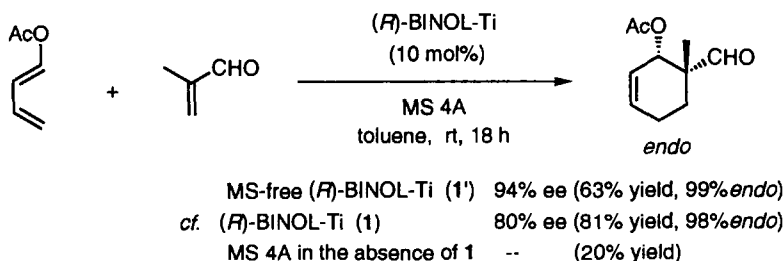


Scheme 6.32

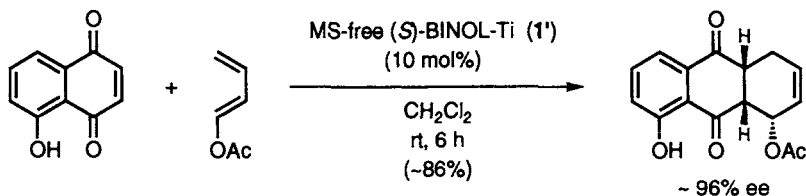
The chiral titanium complex derived from 6-Br-BINOL affords higher enantioselectivity and catalytic activity than the parent BINOL-Ti catalyst in the hetero D-A reactions of 1-methoxydienes with glyoxylate, but not with bromoacrolein (Scheme 6.34).¹⁵⁸



The D-A reaction of methacrolein with 1,3-dienol derivatives can also be catalysed by the BINOL-derived titanium complex. However, the catalyst must be freed from MS to give the *endo*-adduct with high enantioselectivity (Scheme 6.35) (Protocol 5).^{159,160} Because MS works as an achiral catalyst for the D-A reaction. The asymmetric D-A reaction catalysed by the MS-free (MS-(–)) BINOL-Ti complex (**1'**) can be applied to naphthoquinone derivatives as the dienophile to provide an entry to the asymmetric synthesis of tetra- and anthracyclinone^{161–163} aglycones (Scheme 6.36). The sense of asymmetric induction is exactly the same as observed for the asymmetric catalytic reactions described above in the presence of MS.



6: Chiral titanium complexes for enantioselective catalysis

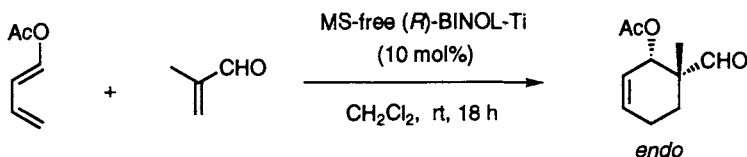


Scheme 6.36

Protocol 5.

Asymmetric Diels–Alder reaction catalyzed by MS-free BINOL-derived titanium complex^{159,160} (Scheme 6.37)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 6.37

Equipment

- A pre-dried, two-necked, round-bottomed flask (100 mL) equipped with a magnetic stirring bar, a three-way stopcock, and a septum cap. The three-way stopcock is connected to an argon source
- A pre-dried, two-necked, round-bottomed flask (100 mL) equipped with a magnetic stirring bar, a three-way stopcock, and a septum cap. The three-way stopcock is connected to a vacuum/argon source
- A pre-dried distillation apparatus
- Two pre-dried centrifuge tubes (40 mL) equipped with a septum cap
- A pre-dried, two-necked, round-bottomed flask (20 mL) equipped with a magnetic stirring bar, a three-way stopcock, and a septum cap. The three-way stopcock is connected to an argon source
- Dry gas-tight syringe
- Double-ended cannula
- Thermometers for low temperature

Materials

- Molecular sieves 4A,^a powder, <5 μm , activated, 10 g
- (*R*)-(+)- or (*S*)-(–)-1,1'-bi-2-naphthol (FW 286.3), 0.573 g, 2.0 mmol
- Di-isopropoxytitanium(IV) dibromide (FW 325.9), 0.652 g, 2.0 mmol
- 1-Acetoxy-1,3-butadiene (mixture of *trans/cis*, 6:4) (FW 112.1), 198 mL, 1.67 mmol (*trans* isomer 1.0 mmol)
- Freshly distilled methacrolein (FW 70.1), 82.8 mL, 1.0 mmol
- Dry dichloromethane
- Dry toluene
- Triethylamine (FW 101.2), 70 mL, 0.5 mmol

irritant, hygroscopic
irritant
moisture sensitive, irritant

flammable liquid, toxic
flammable liquid, corrosive
toxic, irritant
flammable liquid, toxic
flammable liquid, corrosive

1. Add the powdered molecular sieves 4A (10 g) and (*R*)-(+)-1,1'-bi-2-naphthol (0.573 g, 2.00 mmol) to the pre-dried two-necked flask, purge with argon, and add the dry CH_2Cl_2 (60 mL). Stir for 15 min at room temperature.

Protocol 5. Continued

2. Add di-isopropoxytitanium dibromide (0.652 g, 2.0 mmol) (prepared as described in protocol 1) to the suspension in one portion under a flow of argon. At this point the reaction mixture will turn into a red-brown suspension.
3. After stirring for 1 h at room temperature, transfer the suspension with the double-ended cannula to two pre-dried centrifuge tubes capped with a septum by pressurizing with argon.
4. Sediment the molecular sieves 4A by centrifugation at 4000 r.p.m. for 20 min. Transfer the resulting supernatant solvent with the double-ended cannula into the pre-dried two-necked round-bottomed flask. Exchange the septum cap with a distillation apparatus under a flow of argon.
5. Slowly evaporate the solvent *in vacuo* at 0°C. Control the speed of the distillation by regulation of the vacuum.
6. Dry the resulting dark-red solid *in vacuo* and obtain 0.58–0.62 g of the BINOL–Ti complex.^b The MS-free BINOL–Ti complex may be stored under argon in refrigerator.
7. Add the MS-free BINOL–Ti complex (43 mg, 0.1 mmol^b) to the pre-dried two-necked round-bottomed flask. Dissolve the BINOL–Ti complex with toluene (3 mL).
8. Add freshly distilled methacrolein (82.8 mL, 1 mmol) and a solution of 1-acetoxy-1,3-butadiene (198 mL, 1.67 mmol) in toluene (1 mL) at room temperature.
9. After stirring for 18 h at room temperature, dilute the resulting solution with ether (5 mL) and quench with a solution of triethylamine (70 mL, 0.5 mmol) in hexane (10 mL). At this point, a yellow precipitate will form.
10. Filter off the precipitate through a pad of Celite and florisil, and rinse the filter cake three times with ether.
11. Concentrate the crude product *in vacuo* and purify the residue by silica gel chromatography to collect 115 mg (63%) of 2-acetoxy-1-methylcyclohex-3-enecarbaldehyde. The product displays the appropriate spectral characteristics and high resolution mass spectral data.
12. The enantiomeric purity is determined by the capillary GC analysis using a CP-cyclodextrin- β -2,3,6-M-19 as a chiral stationary phase column (94% ee).

^a Purchased from Aldrich Chemical Company, Inc. and used without activation.

^b Exact molecular weight of the MS-free BINOL–Ti complex has not been determined yet. However, on the basis of elemental analysis of the BINOL–Ti complex, there exists almost any bromine (<0.1%) and the content of titanium is ranging from 10.8% to 11.2%. The catalyst mol% is therefore calculated by its titanium content (averaged value 11.0%) wherein 434 mg of the BINOL–Ti complex equal 1 mmol of titanium.

Table 6.5 NLE in asymmetric D–A reaction of 1-acetoxy-1,3-butadiene and methacrolein catalyzed by MS-free BINOL-Ti (**1'**)

| Run | MS-free BINOL-Ti (1') (% ee) | %yield | %endo | %ee |
|----------------|--|--------|-------|-----|
| 1 ^a | 52 | 41 | 98 | 76 |
| 2 ^b | 50 | 50 | 99 | 74 |
| 3 ^c | 50 | 62 | 99 | 40 |
| 4 ^d | 50 | 67 | 99 | 60 |
| 5 ^e | 50 | 62 | 95 | 29 |
| 6 ^f | – | 20 | – | – |
| 7 ^g | 50 | 52 | 99 | 53 |

^a Prepared from partially resolved BINOL (52% ee) and Cl₂Ti(OPrⁱ)₂.^b MS-free (*R*)-**1'** and MS-free (±)-**1'** (1:1).^c MS-free (*R*)-**1'** and MS-free (*S*)-**1'** (3:1).^d Prepared from MS-free (*R*)-**1'** and MS-free (*S*)-**1'** (3:1) in the presence of MS which was filtered prior to the reaction.^e MS-free (*R*)-**1'** and MS-free (±)-**1'** (1:1) in the presence of MS 4A.^f No MS-free catalyst (**1'**) in the presence of MS 4A.^g MS-free (*R*)-**1'** and MS-free (*S*)-**1'** (3:1) in CH₂Cl₂.

The mode of preparation of the MS-free BINOL-Ti catalyst (**1'**) determines the presence or the absence of a non-linear effect (NLE) (Table 6.5, Fig. 6.6). When the MS-free catalyst (**1'**) was prepared from partially resolved BINOL, a (+)-NLE was observed (Run 1). The combined use of an enantiopure (*R*)-**1'** and (±)-**1'** catalysts in a ratio of 1:1 results in a similar (+)-NLE (Run 2). By contrast, mixing enantiopure (*R*)- and (*S*)-**1'** catalysts in a ratio of 3:1 leads to a linearity (no NLE) (Run 3). However, a MS-free catalyst obtained by mixing (*R*)- and (*S*)-**1'** catalysts in the same ratio of 3:1 *in the presence of MS*, which was filtered off prior to the reaction, showed a (+)-NLE (Run 4). These experimental facts can be explained if the complex consists of oligomers which do not interconvert *in the absence of MS* in toluene but do interconvert in dichloromethane (see Run 7 for the 9(+)-NLE in CH₂Cl₂). When the reaction was carried out *in the presence of MS*, however, a (–)-NLE was observed (Run 5), because MS acts as an achiral catalyst for the D–A reaction (Run 6). Moreover, in dichloromethane, the combined use of (*R*)- and (*S*)-**1'** catalysts (3:1), even without prior treatment with MS, exhibited a (+)-NLE (Run 7).

Asymmetric activation of the BINOL-Ti(OPrⁱ)₂ (**2**) by (6-Br)-BINOLs is essential to provide higher levels of enantioselectivity than those attained by the enantiopure BINOL-Ti catalyst (5% ee) in the D–A reaction of glyoxylates with the Danishefsky diene (Scheme 6.38) (Protocol 6)¹¹⁰

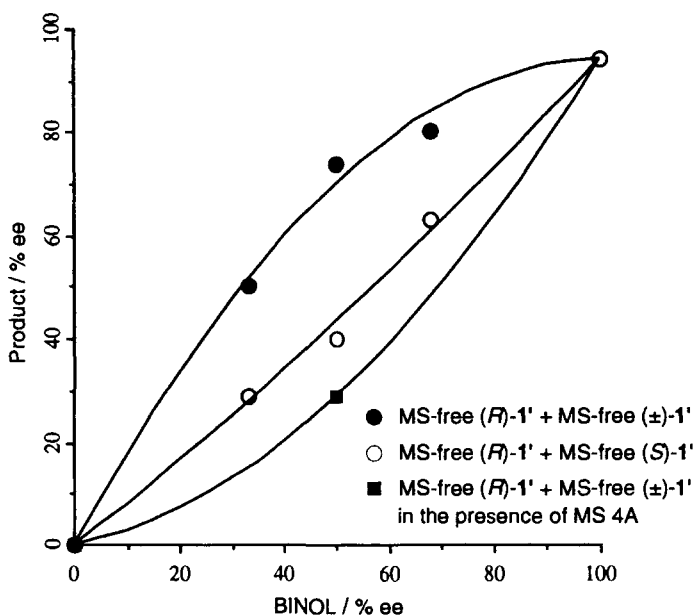
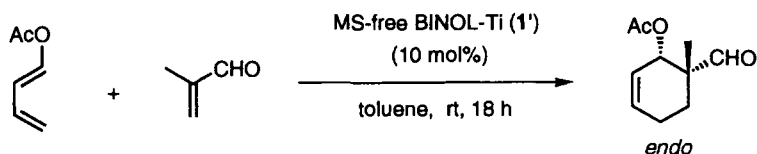
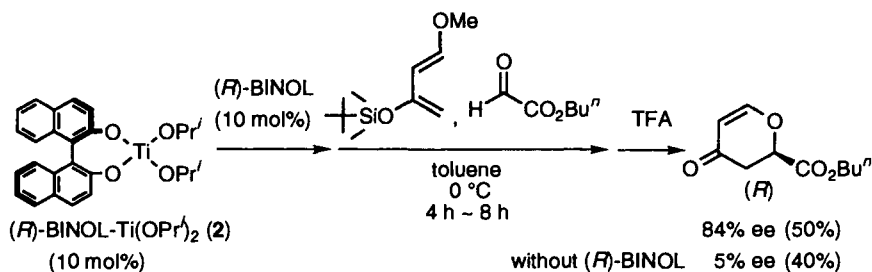


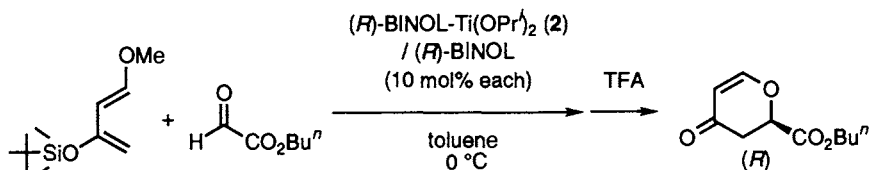
Fig. 6.6. (+)-, (-)-NLE, and linear relationships depending on the catalyst preparation.



Scheme 6.38

Protocol 6.**Asymmetric hetero Diels–Alder reaction catalysed by (*R*)-BINOL-Ti(OPr^{*i*})₂/*(R)*-BINOL complex: Asymmetric activation of (*R*)-BINOL-Ti(OPr^{*i*})₂ by additional (*R*)-BINOL¹¹⁰ (Scheme 6.39)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 6.39****Equipment**

- A pre-dried, two-necked, round-bottomed flask (100 mL) equipped with a magnetic stirring bar, a three-way stopcock, and a septum cap. The three-way stopcock is connected to an argon source
- Dry gas-tight syringe
- Thermometers for low temperature

Materials

- (*R*)-(+)-1,1'-bi-2-naphthol (FW 286.3), 28.6 mg, 0.1 mmol × 2
- Titanium(IV) isopropoxide (FW 284.3), 0.297 mL, 0.1 mmol
- (*E*)-1-Methoxy-3-((*tert*-butyldimethylsilyl)oxy)-1,3-butadiene (FW 214.4), 257 mg, 1.2 mmol
- Freshly distilled *n*-butyl glyoxylate^a (FW 130.1), 130.1 mg, 1.0 mmol
- Dry toluene
- Trifluoroacetic acid (FW 114.0), 0.1 mL, 1.3 mmol

irritant
moisture sensitive, irritant

flammable liquid, toxic
lachrymator, irritant
flammable liquid, toxic
corrosive, toxic

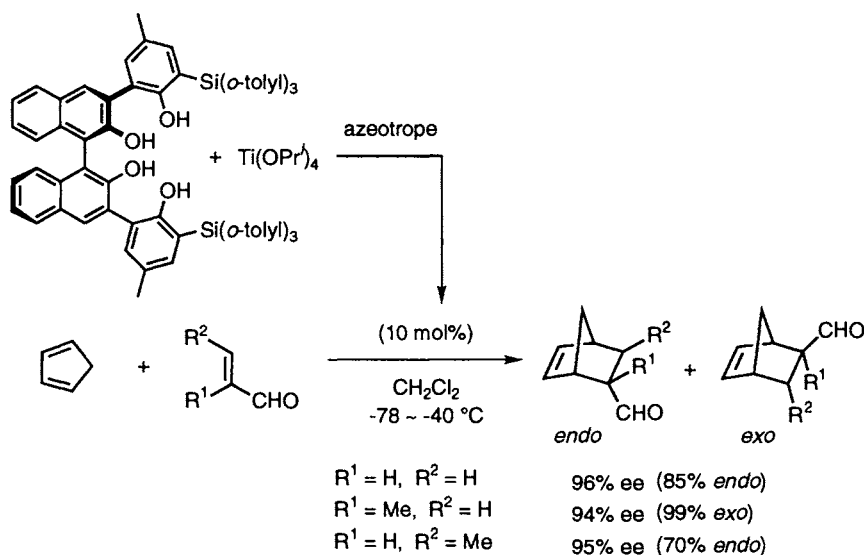
1. Add (*R*)-(+)-1,1'-bi-2-naphthol (28.6 mg, 0.1 mmol) to the pre-dried two-necked flask, purge with argon, and add the dry toluene (2 mL) and the titanium(IV) isopropoxide (0.297 mL, 0.1 mmol) in this order. Stir for at room temperature 20 min. At this point the mixture will turn into a yellow–orange solution.
2. Add additional amount of (*R*)-(+)-1,1'-bi-2-naphthol (28.6 mg, 0.1 mmol) to the solution in one portion under a flow of argon. The mixture will turn immediately into a red–brown solution.
3. After stirring at room temperature for 20 min, cool the catalyst solution to 0 °C. Add the (*E*)-1-methoxy-3-((*tert*-butyldimethylsilyl)oxy)-1,3-butadiene (257 mg, 1.2 mmol) and the freshly distilled *n*-butyl glyoxylate (130.1 mg, 1.0 mmol) in this order to the catalyst solution.
4. After stirring at 0 °C for 2 h, add the trifluoroacetic acid (0.1 mL, 1.3 mmol) to the reaction mixture at that temperature.

Protocol 6. Continued

5. Stir for an additional 5 min at 0°C and then add saturated NaHCO₃ (10 mL) at that temperature. Filter the resulting yellowish suspension through a pad of Celite and rinse the filter cake with ether.
6. Separate the phases of the filtrate and extract the aqueous phase three times with ether (15 mL). Wash the combined organic phases twice with brine, dry over magnesium sulfate, and filter the solution.
7. Concentrate the crude product under vacuum and purify the crude product by silica-gel column chromatography with hexane–ethyl acetate (10:1) as eluent to obtain butyl 3,4-dihydro-4-oxo-2*H*-pyran-2-carboxylate (99.1 mg, 50% yield).
8. The product displays the appropriate spectral characteristics. The enantiomeric purity is determined by the HPLC analysis using a CHIRALPAK AS as a chiral stationary phase column with 10% *i*-PrOH/hexane as a mobile phase (84 %ee).

^a Immediately prior use, the *n*-butyl glyoxylate is depolymerized by vacuum distillation from phosphorus pentoxide (10% weight) bp. 80°C/35 mmHg, bath temperature 120–140°C.

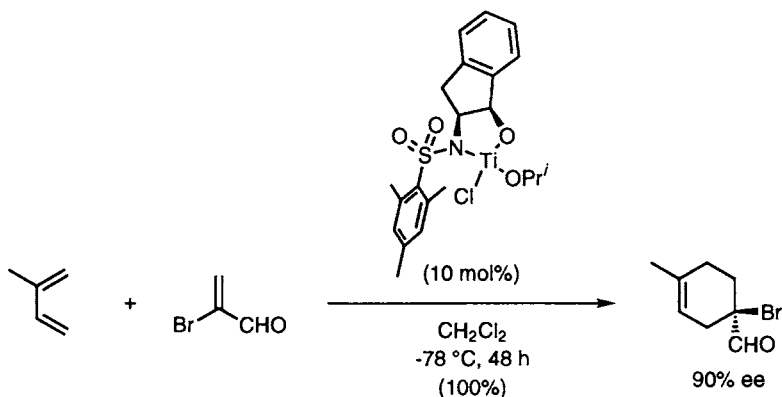
As shown above, asymmetric catalysis of the D–A reactions have been attained using chiral titanium complexes bearing chiral diol ligands.^{164–166} Whereas Yamamoto has reported a chiral helical titanium complex derived from Ti(OPr^{*i*})₄ and a BINOL-derived tetraol ligand (Scheme 6.40).¹⁶⁷ The

**Scheme 6.39**

6: Chiral titanium complexes for enantioselective catalysis

D–A products are obtained with uniformly high enantioselectivity, irrespective of the substituent pattern of α,β -unsaturated aldehydes.

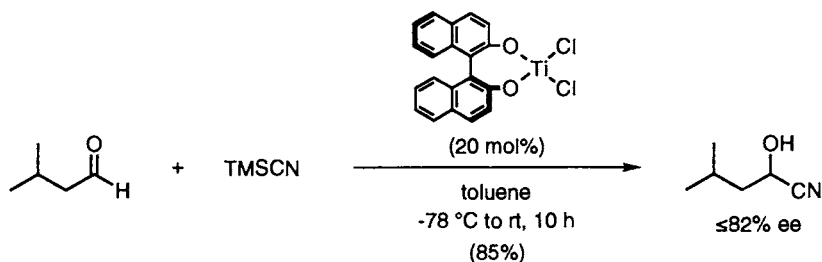
Corey has also reported a new type of chiral titanium complex, which is derived from an amino alcohol ligand¹⁶⁸ (Scheme 6.41). The chiral titanium complex serves as an efficient asymmetric catalyst for the reaction of 2-bromoacrolein to yield the D–A product in high enantioselectivity.



Scheme 6.41

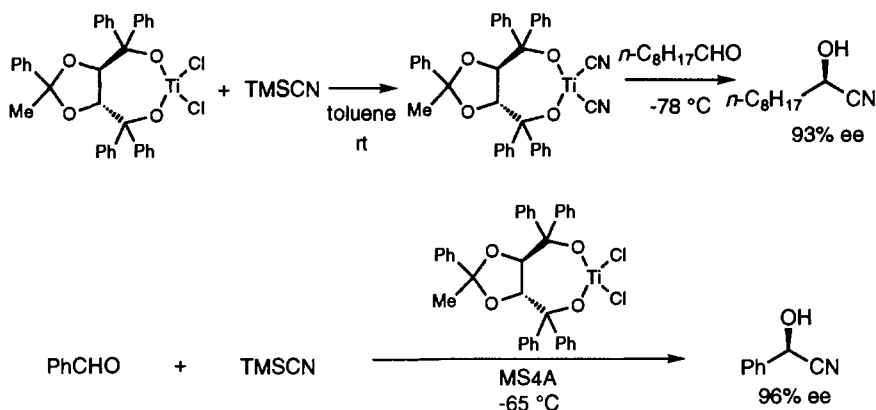
11. Cyanohydrine formation

In the addition reaction of cyanotrimethylsilane¹⁶⁹ to aliphatic aldehydes, another synthetic application of a BINOL-Ti catalyst was reported by Reetz.¹²² In his case, however, BINOL- TiCl_2 was reported to be prepared by treatment of the lithium salt of BINOL with TiCl_4 in ether (see above). The BINOL- TiCl_2 thus obtained was used as a catalyst for the cyanosilylation reaction to give the cyanohydrins in $<82\%$ ee (Scheme 6.42).



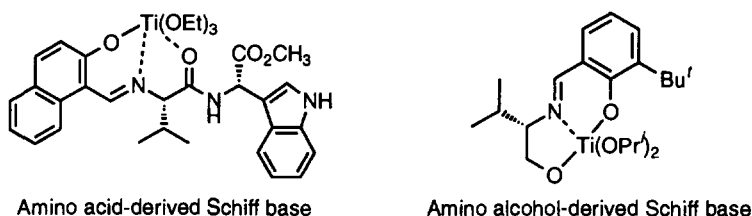
Scheme 6.42

Narasaka has also reported that TADDOL-Ti dichloride acts as a good catalyst for the asymmetric addition of trimethylsilylcyanide to aromatic and aliphatic aldehydes (Scheme 6.43).¹⁷⁰ The reactions proceeded only in the



Scheme 6.43

presence of MS 4A. In the reaction with aromatic aldehydes, a chiral cyano-titanium species, which is obtained by mixing of the TADDOL-Ti dichloride and trimethylsilylcyanide prior to addition of the aldehydes, acts as a better chiral cyanating agent to afford higher enantiomeric excesses. Chiral titanium complexes obtained after addition of a salicylaldehyde-type Schiff bases have been reported to catalyse the asymmetric addition of hydrogen cyanide¹⁷¹ or trimethylsilylcyanide^{172–174} to aromatic and aliphatic aldehydes with high enantioselectivity (Scheme 6.44).



Scheme 6.44

12. Miscellaneous reactions

Chiral titanium complexes are also used as effective asymmetric catalysts for other carbon–carbon bond forming reactions, such as inverse electron-demand Diels–Alder reactions,^{175–177} [2+2] additions,^{178,179} [2+3] additions,¹⁸⁰ Michael additions,^{181,182} and others^{183–187}. From a practical point of view, the development of more active and efficient catalysts is important, for the molecular design of asymmetric catalysts is the key to the basis of the structure–catalytic activity relationship. Although structure determination of active titanium species has been quite limited so far,^{188–194} any progress along this line is highly promising and worth the effort.

Acknowledgements

We are grateful to Drs Darren Smyth and Thorsten Volk for their eagle-eyed proofreading of the manuscript.

References

1. Noyori, R. *Science* **1990**, *248*, 1194–1199.
2. Noyori, R. *Asymmetric Catalysis in Organic Synthesis*. Wiley, New York, **1994**.
3. Kagan, H. B. *Asymmetric Synthesis*. Georg Thieme Verlag, Stuttgart, **1997**.
4. Reetz, M. T. *Organotitanium Reagents in Organic Synthesis*. Springer-Verlag, Berlin, **1986**.
5. Seebach, D. *Angew. Chem., Int. Ed. Engl.* **1990**, *29*, 1320–1367.
6. Brown, I. D. *Acta Cryst.* **1988**, *B44*, 545–553.
7. Duthaler, R. O.; Hafner, A.; Riediker, M. *Organic Synthesis via Organometallics*. Friedr. Vieweg, Braunschweig, **1991**, p 285–309.
8. Duthaler, R. O.; Hafner, A.; Riediker, M. *Pure Appl. Chem.* **1990**, *62*, 631–642.
9. Mikami, K.; Terada, M.; Nakai, T. *Kikan Kagaku Sosetsu No. 17: Organic Chemistry of the Early Transition Metals*. Gakkai Shuppan Center, Tokyo, **1993**, p 87–98.
10. Mikami, K.; Nakai, T. *Kagaku Zoukann No. 124*. Kagaku Doujinn, Kyoto, **1995**, p 177–192.
11. Duthaler, R. O.; Hafner, A. *Chem. Rev.* **1992**, *92*, 807–832.
12. Bradley, D. C.; Mehrotra, R. C.; Gaur, D. P. *Metal Alkoxides*. Academic Press, New York, **1978**.
13. Feld, R.; Cowe, P. L. *The Organic Chemistry of Titanium*. Butterworths, London, **1965**.
14. Katsuki, T.; Martin, V. S. *Org. React.* **1996**, *48*, 1–299.
15. Johnson, R. A.; Sharpless, K. B. In *Comprehensive Organic Synthesis* (ed. B. M. Trost, I. Fleming), Pergamon Press, Oxford, **1991**, Vol. 7, pp 389–436.
16. Finn, M. G.; Sharpless, K. B. In *Asymmetric Synthesis* (ed. J. D. Morrison) Academic Press, New York, **1985**, Vol. 5, pp 247–308.
17. Rossiter, B. E. In *Asymmetric Synthesis* (ed. J. D. Morrison). Academic Press, New York, **1985**, Vol. 5, pp 193–246.
18. Halterman, R. L. *Chem. Rev.* **1992**, *92*, 965–994.
19. Carter, M. B.; Schiott, B.; Gutierrez, A.; Buchwald, S. L. *J. Am. Chem. Soc.* **1994**, *116*, 11667–11670.
20. Willoughby, C. A.; Buchwald, S. L. *J. Am. Chem. Soc.* **1994**, *116*, 11703–11714.
21. Almquist, F.; Torsysson, L.; Gudmundsson, A.; Frejd, T. *Angew. Chem., Int. Ed. Engl.* **1997**, *36*, 367–377.
22. Lindsley, C. W.; DiMare, M. *Tetrahedron Lett.* **1994**, *35*, 5141–5144.
23. Giffels, G.; Dreisbach, C.; Kragl, U.; Wegerding, M.; Waldmann, H.; Wandrey, C. *Angew. Chem., Int. Ed. Engl.* **1995**, *34*, 2005–2006.
24. Mikami, K.; Terada, M.; Narisawa, S.; Nakai, T. *Org. Synth.* **1992**, *71*, 14–21.
25. Dijkgraaf, C.; Rousseau, J. P. G. *Spectrochim. Acta* **1968**, *2*, 1213–1217.
26. Reetz, M. T.; Kyung, S.-H.; Westermann, J. *Organometallics* **1984**, *3*, 1716–1717.
27. Mikami, K.; Motoyama, Y. In *Encyclopedia of Reagents for Organic Synthesis* (ed. L. A. Paquette), Wiley, Chichester, **1995**, Vol. 1, pp 397–403.

28. Mikami, K. In *Encyclopedia of Reagents for Organic Synthesis* (ed. L. A. Paquette), Wiley, Chichester, **1995**, Vol. 1, pp 407–408.
29. Wang, J.-T.; Fan, X.; Feng, X.; Qian, Y.-M. *Synthesis* **1989**, 291–292.
30. Olivero, A. G.; Weidmann, B.; Seebach, D. *Helv. Chim. Acta* **1982**, 64, 2485–2488.
31. Braun, M. *Angew. Chem., Int. Ed. Engl.* **1996**, 35, 519–522.
32. Dahinden, R.; Beck, A. K.; Seebach, D. In *Encyclopedia of Reagents for Organic Synthesis* (ed. L. A. Paquette), Wiley, Chichester, **1995**, Vol. 3, pp 2167–2170.
33. Narasaka, K.; Iwasawa, N.; Inoue, M.; Yamada, T.; Nakashima, M.; Sugimori, J. *J. Am. Chem. Soc.* **1989**, 111, 5340–5344.
34. Beck, A. K.; Bastani, B.; Plattner, D. A.; Petter, W.; Seebach, D. *Chimia* **1991**, 45, 238–244.
35. Hafner, A.; Duthaler, R. O.; Marti, R.; Rihs, G.; Rothe-Streit, P.; Schwarzenbach, F. *J. Am. Chem. Soc.* **1992**, 114, 2321–2336.
36. Seebach, D.; Behrendt, L.; Felix, D. *Angew. Chem., Int. Ed. Engl.* **1991**, 30, 1008–1009.
37. Bussche-Hünnefeld, J. L.; Seebach, D. *Tetrahedron* **1992**, 48, 5719–5730.
38. Takahashi, H.; Kawakita, T.; Ohno, M.; Yoshioka, M.; Kobayashi, S. *Tetrahedron* **1992**, 48, 5691–5700.
39. Yoshioka, M.; Kawakita, T.; Ohno, M. *Tetrahedron Lett.* **1989**, 30, 1657–1660.
40. Takahashi, H.; Kawakita, T.; Yoshioka, M.; Kobayashi, S.; Ohno, M. *Tetrahedron Lett.* **1989**, 30, 7095–7098.
41. Eisenberg, C.; Knochel, P. *J. Org. Chem.* **1994**, 59, 3760–3761.
42. Rozema, M. J.; Sidduri, A.; Knochel, P. *J. Org. Chem.* **1992**, 57, 1956–1958.
43. Knochel, P. In *Comprehensive Organic Synthesis* (eds. B. M. Trost, I. Fleming). Pergamon, London, **1991**, Vol. 1, pp 211–229.
44. Mikami, K. In *Encyclopedia of Reagents for Organic Synthesis* (ed. L. A. Paquette), Wiley, Chichester, **1995**, Vol. 1, pp 403–406.
45. Aoki, S.; Mikami, K.; Terada, M.; Nakai, T. *Tetrahedron* **1993**, 49, 1783–1792.
46. Roush, W. R.; Blizzad, T. A. *Tetrahedron Lett.* **1982**, 23, 2331–2334.
47. Still, W. C.; Ohmizu, H. *J. Org. Chem.* **1981**, 46, 5242–5244.
48. Mikami, K.; Matsukawa, S. *Tetrahedron Lett.* **1994**, 35, 3133–3166.
49. Costa, A. L.; Piazza, M. G.; Tagliavini, E.; Trombini, C.; Umani-Ronchi, A. *J. Am. Chem. Soc.* **1993**, 115, 7001–7002.
50. Keck, G. E.; Tarbet, K. H.; Geraci, L. S. *J. Am. Chem. Soc.* **1993**, 115, 8467–8468.
51. Keck, G. E.; Krishnamurthy, D.; Grier, M. C. *J. Org. Chem.* **1993**, 58, 6543–6544.
52. Weigand, S.; Brückner, R. *Chem. Eur. J.* **1996**, 2, 1077–1084.
53. Gauthier, Jr, D. R.; Carreira, E. M. *Angew. Chem., Int. Ed. Engl.* **1996**, 35, 2363–2365.
54. Mikami, K.; Shimizu, M. *Chem. Rev.* **1992**, 92, 1021–1050.
55. Snider, B. B. In *Comprehensive Organic Synthesis* (ed. B. M. Trost, I. Fleming). Pergamon, London, **1991**, Vol. 2, pp 527–561, Vol. 5, pp 1–27.
56. Mikami, K.; Terada, M.; Shimizu, M.; Nakai, T. *J. Synth. Org. Chem. Jpn* **1990**, 48, 292–303.
57. Hoffmann, H. M. R. *Angew. Chem., Int. Ed. Engl.* **1969**, 8, 556–577.
58. Mikami, K. *Advances in Asymmetric Synthesis*. JAI Press, Greenwich, Connecticut, **1995**, Vol. 1, pp 1–44.
59. Mikami, K.; Loh, T.-P.; Nakai, T. *Tetrahedron Lett.* **1988**, 29, 6305–6308.
60. Mikami, K.; Loh, T.-P.; Nakai, T. *J. Chem. Soc., Chem. Comm.* **1988**, 1430–1431.

61. Omura, S. *J. Synth. Org. Chem. Jpn.* **1986**, 44, 127.
62. Hanessian, S. *Total Synthesis of Natural Products: The 'Chiron' Approach*. Pergamon, Oxford, **1983**.
63. Mori, K. *The Total Synthesis of Natural Products*. Wiley, New York, **1981**, Vol. 4.
64. Seebach, D.; Hungerbughler, E. *Modern Synthetic Methods* (ed. R. Scheffold). Otto Salle Verlag, Frankfurt am Main, **1980**, Vol. 2, pp 91–172.
65. Mikami, K. *Pure Appl. Chem.* **1996**, 68, 639–644.
66. Mikami, K.; Terada, M.; Nakai, T. In *Advances in Catalytic Processes*, (ed. M. P. Doyle). JAI Press, London, **1995**, Vol. 1, p 123–149.
67. Mikami, K.; Terada, M.; Narisawa, S.; Nakai, T. *Synlett* **1992**, 255–265.
68. Mikami, K.; Terada, M.; Nakai, T. *J. Am. Chem. Soc.* **1989**, 111, 1940–1941.
69. Mikami, K.; Terada, M.; Nakai, T. *J. Am. Chem. Soc.* **1990**, 112, 3949–3954.
70. Maruoka, K.; Hoshino, Y.; Shirasaka, T.; Yamamoto, H. *Tetrahedron Lett.* **1988**, 29, 3967–3970.
71. Mikami, K.; Terada, M.; Nakai, T. *Chem. Express* **1989**, 4, 589–592.
72. Mikami, K.; Motoyama, Y.; Terada, M. *Inorg. Chim. Acta* **1994**, 222, 71–75.
73. Terada, M.; Motoyama, Y.; Mikami, K. *Tetrahedron Lett.* **1994**, 35, 6693–6696.
74. Terada, M.; Mikami, K. *J. Chem. Soc. Chem. Commun.* **1995**, 2391–2392.
75. Terada, M.; Matsukawa, S.; Mikami, K. *J. Chem. Soc., Chem. Commun.* **1993**, 327–328.
76. Mori, K.; Takigawa, H. *Tetrahedron* **1991**, 47, 2163–2168.
77. Brown, H. C.; Randad, R. S. *Tetrahedron* **1990**, 46, 4463–4472.
78. Ohloff, G.; Giersch, W. *Helv. Chim. Acta* **1977**, 60, 1496–1500.
79. Mikami, K.; Yajima, T.; Terada, M.; Uchimar, T. *Tetrahedron Lett.* **1993**, 34, 7591–7594.
80. Mikami, K.; Yajima, T.; Terada, M.; Kato, E.; Maruta, M. *Tetrahedron Asymm.* **1994**, 5, 1087–1090.
81. Mikami, K.; Yajima, T.; Takasaki, T.; Matsukawa, S.; Terada, M.; Uchimar, T.; Maruta, M. *Tetrahedron* **1996**, 52, 85–98.
82. Welch, J. T.; Eswarakrishnan, S. *Flurine in Bioorganic Chemistry*. Wiley, New York, **1990**.
83. Mikami, K.; Siree, N.; Yajima, T.; Terada, M.; Suzuki, Y. Annual Meeting of the Chemical Society of Japan, Kyoto, March 28–31, 1995; Abstract No. 3H218.
84. Mikami, K.; Yajima, T.; Siree, N.; Terada, M.; Suzuki, Y.; Kobayashi, I. *Synlett* **1996**, 837–838.
85. Mikami, K.; Yajima, T.; Terada, M.; Kawauchi, S.; Suzuki, Y.; Kobayashi, I. *Chem. Lett.* **1996**, 861–862.
86. Mikami, K.; Yajima, T.; Terada, M.; Suzuki, Y.; Kobayashi, I. *Chem. Commun.* **1997**, 57–58.
87. Mikami, K.; Yoshida, A. *Synlett* **1995**, 29–31.
88. Mikami, K.; Yoshida, A.; Matsumoto, Y. *Tetrahedron Lett.* **1996**, 37, 8515–8518.
89. Ward, R. S. *Chem. Soc. Rev.* **1990**, 19, 1–19.
90. Mikami, K.; Shimizu, M. *J. Synth. Org. Chem. Jpn.* **1993**, 21–31.
91. Bartlett, P. A. *Tetrahedron* **1980**, 36, 2–72.
92. Mikami, K.; Narisawa, S.; Shimizu, M.; Terada, M. *J. Am. Chem. Soc.* **1992**, 114, 6566–6568; 9242–9242.
93. Kagan, H. B.; Fiaud, J. C. *Topics in Stereochemistry Interscience*, New York, **1988**, Vol. 18.

94. Brown, J. M. *Chem. Ind. (London)* **1988**, 612–617.
95. Masamune, S.; Choy, W.; Peterson, J.; Sita, L. R. *Angew. Chem., Int. Ed. Engl.* **1985**, *24*, 1–30.
96. Heathcock, C. H. In *Asymmetric Synthesis* (ed. J. D. Morrison). Academic Press: New York, **1985**, Vol. 3, pp 111–212.
97. Guillaneux, D.; Zhao, S.-H.; Samuel, O.; Rainford, D.; Kagan, H. B. *J. Am. Chem. Soc.* **1994**, *116*, 9430–9439.
98. Puchot, C.; Samuel, O.; Dunach, E.; Zhao, S.; Agami, C.; Kagan, H. B. *J. Am. Chem. Soc.* **1986**, *108*, 2353–2357.
99. Mikami, K.; Terada, M. *Tetrahedron* **1992**, *48*, 5671–5680.
100. Terada, M.; Mikami, K.; Nakai, T. *J. Chem. Soc., Chem. Commun.* **1990**, 1623–1624.
101. Oguni, N.; Matsuda, Y.; Kaneko, T. *J. Am. Chem. Soc.* **1988**, *110*, 7877–7877.
102. Noyori, R.; Kitamura, M. *Angew. Chem. Int. Ed. Engl.* **1991**, *30*, 49–69.
103. Kitamura, M.; Suga, S.; Niwa, M.; Noyori, R. *J. Am. Chem. Soc.* **1995**, *117*, 4832–4842.
104. Alcock, N. W.; Brown, J. M.; Maddox, P. J. *J. Chem. Soc., Chem. Commun.* **1986**, 1532–1533.
105. Maruoka, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1988**, *111*, 789–790.
106. Faller, J. W.; Parr, J. J. *J. Am. Chem. Soc.* **1993**, *115*, 804–805.
107. Mikami, K.; Matsukawa, S. *Nature* **1997**, *385*, 613–615.
108. Matsukawa, S.; Mikami, K. *Tetrahedron Asymmetry* **1995**, *6*, 2571–2574.
109. Matsukawa, S.; Mikami, K. *Enantiomer* **1996**, *1*, 69–73.
110. Matsukawa, S.; Mikami, K. *Tetrahedron Asymmetry* **1997**, *8*, 815–816.
111. Taber, D. F. *Intramolecular Diels–Alder and Alder Ene Reactions*. Springer Verlag, Berlin, **1984**.
112. Fujita, Y.; Suzuki, S.; Kanehira, K. *J. Synth. Org. Chem. Jpn.* **1983**, *41*, 1152–1167.
113. Oppolzer, W.; Snieckus, V. *Angew. Chem. Int. Ed. Engl.* **1978**, *17*, 476–486.
114. Conia, J. M.; Le Perche, P. *Synthesis* **1975**, 1–19.
115. Mikami, K.; Sawa, E.; Terada, M. *Tetrahedron Asymmetry* **1991**, *2*, 1403–1412.
116. Sakane, S.; Maruoka, K.; Yamamoto, H. *Tetrahedron* **1986**, *42*, 2203–2209.
117. Sakane, S.; Maruoka, K.; Yamamoto, H. *Tetrahedron Lett.* **1985**, *26*, 5535–5538.
118. Narasaka, K.; Hayashi, Y.; Shimada, S. *Chem. Lett.* **1988**, 1609–1612.
119. Mikami, K.; Terada, M.; Sawa, E.; Nakai, T. *Tetrahedron Lett.* **1991**, *32*, 6571–6574.
120. Evans, D. A.; Nelson, J. V.; Taber, T. R. *Topics in Stereochemistry*. Vol. 13, Interscience, New York, **1982**.
121. Mukaiyama, T. *Org. React.* **1982**, *28*, 203–331.
122. Reetz, M. T.; Kyung, S.-H.; Bolm, C.; Zierke, T. *Chem. Ind. (London)* **1986**, 824–824.
123. Mukaiyama, T.; Inubushi, A.; Suda, S.; Hara, R.; Kobayashi, S. *Chem. Lett.* **1990**, 1015–1018.
124. Kitamoto, D.; Imma, H.; Nakai, T. *Tetrahedron Lett.* **1995**, *36*, 1861–1864.
125. Mikami, K.; Matsukawa, S. *J. Am. Chem. Soc.* **1993**, *115*, 7039–7040.
126. Murata, S.; Suzuki, M.; Noyori, R. *J. Am. Chem. Soc.* **1980**, *102*, 3248–3249.
127. Yamamoto, Y.; Maruyama, K. *Tetrahedron Lett.* **1980**, *21*, 4607–4610.
128. Mikami, K.; Matsukawa, S.; Nagashima, M.; Funabashi, H.; Morishima, H. *Tetrahedron Lett.* **1997**, *38*, 579–582.

129. Mikami, K.; Matsukawa, S.; Sawa, E.; Harada, A.; Koga, N. *Tetrahedron Lett.* **1997**, 38, 1951–1954.
130. Mikami, K.; Matsukawa, S. *J. Am. Chem. Soc.* **1994**, 116, 4077–4078.
131. Kolb, H. C.; Bennari, Y. L.; Sharpless, K. B. *Tetrahedron Asymmetry* **1993**, 4, 133–141.
132. Larcheveque, M.; Henrot, S. *Tetrahedron* **1990**, 46, 4277–4282.
133. Reetz, M. T.; Jung, A. *J. Am. Chem. Soc.* **1983**, 105, 4833–4835.
134. Evans, D. A.; Duffy, J. L.; Dart, M. J. *Tetrahedron Lett.* **1994**, 35, 8537–8540.
135. Evans, D. A.; Dart, M. J.; Duffy, J. L.; Yang, M. G.; Livingston, A. B. *J. Am. Chem. Soc.* **1995**, 117, 6619–6629.
136. Evans, D. A.; Dart, M. J.; Duffy, J. L.; Yang, M. G. *J. Am. Chem. Soc.* **1996**, 118, 4322–4343.
137. Matsukawa, S.; Mikami, K. Annual Meeting of the Chemical Society of Japan, Kyoto, March 28–31, 1995; Abstract No. 3H106.
138. Mikami, K.; Yajima, T.; Takasaki, T.; Matsukawa, S.; Terada, M.; Uchimaru, T.; Maruta, M. *Tetrahedron* **1996**, 52, 85–98.
139. Mikami, K.; Takasaki, T.; Matsukawa, S.; Maruta, M. *Synlett* **1995**, 1057–1058.
140. Pommier, A.; Pons, J.-M. *Synthesis* **1995**, 729–744.
141. Keck, G. E.; Krishnamurthy, D. *J. Am. Chem. Soc.* **1995**, 117, 2363–2364.
142. Carreira, E. M.; Singer, R. A.; Lee, W. *J. Am. Chem. Soc.* **1994**, 116, 8837–8838.
143. Carreira, E. M.; Lee, W.; Singer, R. A. *J. Am. Chem. Soc.* **1995**, 117, 3649–3650.
144. Keck, G. E.; Li, X.-Y.; Krishnamurthy, D. *J. Org. Chem.* **1995**, 60, 5998–5999.
145. Review: Kagan, H. B.; Riant, O. *Chem. Rev.* **1992**, 92, 1007–1019.
146. Narasaka, K. *Synthesis* **1991**, 1–11.
147. Narasaka, K.; Iwasawa, N.; Inoue, M.; Yamada, T.; Nakashima, M.; Sugimori, J. *J. Am. Chem. Soc.* **1989**, 111, 5340–5345.
148. Narasaka, K.; Tanaka, H.; Kanai, F. *Bull. Chem. Soc. Jpn.* **1991**, 64, 387–391.
149. Corey, E. J.; Matsumura, Y. *Tetrahedron Lett.* **1991**, 32, 6289–6292.
150. Seebach, D.; Beck, A. K.; Imwinkelried, R.; Roggo, S.; Wonnacott, A. *Helv. Chim. Acta* **1987**, 70, 954–975.
151. Engler, T. A.; Letavic, M. A.; Takusagawa, F. *Tetrahedron Lett.* **1992**, 33, 6731–6734.
152. Wada, E.; Pei, W.; Kanemasa, S. *Chem. Lett.* **1994**, 2345–2348.
153. Quinkert, G.; Grosso, M. D.; Döring, A.; Döring, W.; Schenkel, R. I.; Bauch, M.; Dambacher, G. T.; Bats, J. W.; Zimmermann, G.; Dürner, G. *Helv. Chim. Acta* **1995**, 78, 1345–1391.
154. Terada, M.; Mikami, K.; Nakai, T. *Tetrahedron Lett.* **1991**, 32, 935–938.
155. Konowal, A.; Jurczak, J.; Zamojski, A. *Tetrahedron* **1976**, 32, 2957–2959.
156. Danishefsky, S. J.; DeNinno, M. P. *Angew. Chem., Int. Ed. Engl.* **1987**, 26, 15–23.
157. Rosen, T.; Heathcock, C. H. *Tetrahedron* **1986**, 42, 4909–4951.
158. Motoyama, Y.; Terada, M.; Mikami, K. *Synlett* **1995**, 967–968.
159. Mikami, K.; Motoyama, Y.; Terada, M. *J. Am. Chem. Soc.* **1994**, 116, 2812–2820.
160. Mikami, K.; Terada, M.; Motoyama, Y.; Nakai, T. *Tetrahedron Asymmetry* **1991**, 2, 643–646.
161. Krohn, K. *Tetrahedron* **1990**, 46, 291–318.
162. Krohn, K. *Angew. Chem., Int. Ed. Engl.* **1986**, 25, 790–807.
163. Broadhurst, M. J.; Hassall, C. H.; Thomas, G. J. *Chem. Ind. (London)* **1985**, 18, 106–112.

164. Devine, P. N.; Oh, T. *J. Org. Chem.* **1992**, *57*, 396–399.
165. Ketter, A.; Glahsl, G.; Herrmann, R. *J. Chem. Research (M)* **1990**, 2118–2156.
166. Chapuis, C.; Jurczak, J. *Helv. Chim. Acta* **1987**, *70*, 436–440.
167. Maruoka, K.; Murase, N.; Yamamoto, H. *J. Org. Chem.* **1993**, *58*, 2938–2939.
168. Corey, E. J.; Roper, T. D.; Ishihara, K.; Sarakinos, G. *Tetrahedron Lett.* **1993**, *34*, 8399–8402.
169. Rasmussen, J. K.; Heilmann, S. M.; Krepski, L. R. *The Chemistry of Cyano-trimethylsilane, Advances in Silicon Chemistry*, JAI Press, Greenwich, **1991**, Vol. 1.
170. Minamikawa, H.; Hayakawa, S.; Yamada, T.; Iwasawa, N.; Narasaka, K. *Bull. Chem. Soc. Jpn.*, **1988**, *61*, 4379–4383.
171. Nitta, H.; Yu, D.; Kudo, M.; Mori, A.; Inoue, S. *J. Am. Chem. Soc.* **1992**, *114*, 7969–7975.
172. Hayashi, M.; Miyamoto, Y.; Inoue, T.; Oguni, N. *J. Org. Chem.* **1993**, *58*, 1515–1522.
173. Hayashi, M.; Inoue, T.; Miyamoto, Y.; Oguni, N. *Tetrahedron* **1994**, *50*, 4385–4398.
174. Hayashi, M.; Matsuda, T.; Oguni, N. *J. Chem. Soc., Perkin Trans. 1* **1992**, 3135–3140.
175. Posner, G. H.; Eydoux, F.; Lee, J.-K.; Bull, D. S. *Tetrahedron Lett.* **1994**, *35*, 7541–7544.
176. Posner, G. H.; Dai, H.; Bull, D. S.; Lee, J.-K.; Eydoux, F.; Ishihara, Y.; Welsh, W.; Pryor, N.; Petr, Jr. S. *J. Org. Chem.* **1996**, *61*, 671–676.
177. Wada, E.; Yasuoka, H.; Kanemasa, S. *Chem. Lett.* **1994**, 1637–1640.
178. Engler, T. A.; Letavic, M. A.; Reddy, J. P. *J. Am. Chem. Soc.* **1991**, *113*, 5068–5070.
179. Narasaka, K.; Hayashi, Y.; Shimadzu, H.; Niihata, S. *J. Am. Chem. Soc.* **1992**, *114*, 8869–8885.
180. Gothelf, K. V.; Thomsen, I.; Jørgensen, K. A. *J. Am. Chem. Soc.* **1996**, *118*, 59–64.
181. Kobayashi, S.; Suda, S.; Yamada, M.; Mukaiyama, T. *Chem. Lett.* **1994**, 97–100.
182. Bernardi, A.; Karamfilova, K.; Boschini, G.; Scolastico, C. *Tetrahedron Lett.* **1995**, *36*, 1363–1366.
183. Inoue, T.; Kitagawa, O.; Oda, Y.; Taguchi, T. *J. Org. Chem.* **1996**, *61*, 8256–8263.
184. Inoue, T.; Kitagawa, O.; Ochiai, O.; Shiro, M.; Taguchi, T. *Tetrahedron Lett.* **1995**, *36*, 9333–9336.
185. Hayashi, M.; Inoue, T.; Oguni, N. *J. Chem. Soc., Chem. Commun.* **1994**, 341–342.
186. Corey, E. J.; Rao, S. A.; Noe, M. C. *J. Am. Chem. Soc.* **1994**, *116*, 9345–9346.
187. Yamazaki, S.; Tanaka, M.; Yamabe, S. *J. Org. Chem.* **1996**, *61*, 4046–4050.
188. Williams, I. D.; Pedersen, S. F.; Sharpless, K. B.; Lippard, S. L. *J. Am. Chem. Soc.* **1984**, *106*, 6430–6431.
189. Bachand, B. Wuest, J. D. *Organometallics* **1991**, *10*, 2015–2025.
190. Boyle, T. J.; Eilerts, N. W.; Heppert, J. A.; Takusagawa, F. *Organometallics* **1994**, *13*, 2218–2229.
191. Corey, E. J.; Letavic, M. A.; Noe, M. C.; Sarshar, S. *Tetrahedron Lett.* **1994**, *35*, 7553–7556.
192. Nugent, W. A.; Harlow, R. L. *J. Am. Chem. Soc.* **1994**, *116*, 6142–6148.
193. Gothelf, K. V.; Hazell, R. G.; Jørgensen, K. A. *J. Am. Chem. Soc.* **1995**, *117*, 4435–4436.
194. Terada, M.; Matsumoto, Y.; Nakamura, Y.; Mikami, K. *Chem. Commun.* **1997**, 281–282.

Tin Lewis acid

SHŪ KOBAYASHI

1. Introduction

The element tin has played an increasingly important role in organic chemistry as well as organometallic chemistry, serving as a source of new reagents for selective transformations.¹⁻³ The main activity in these fields has been focused for a long time on tin(IV) compounds, and tin(II) compounds have been used primarily as reductants of aromatic nitro compounds to aromatic amines.⁴ During the last decade, asymmetric synthesis has been developed increasingly, and in this field tin(II) reagents have served main roles rather than tin(IV) reagents.

Electronegativity of tin(II) and tin(IV) is shown in Table 7.1.⁵ Tin(II) is more electropositive and hence cationic than tin(IV), and is expected to coordinate with nucleophilic ligands. On the other hand, covalent radius and ionic radius of tin(II) are longer than those of tin(IV) (Tables 7.2 and 7.3).⁶⁻¹³ This is due to electronic repulsion caused by unpaired electrons of tin(II), which weakens the δ -bond because tin(II) uses p-orbital for bonding. These

Table 7.1. Comparison of electronegativity

| Oxidation state | Electronegativity (Pauling) | Electronegativity (Sanderson) |
|-----------------|--------------------------------|----------------------------------|
| Sn(II) | 1.80 | 1.58 |
| Sn(IV) | 1.96 | 2.02 |

Table 7.2. Sn-X bond distance (Å) in gas-phase

| SnX ₂ | Sn-X | SnX ₄ | Sn-X | References |
|-------------------|-----------|-------------------|-----------|------------|
| SnR ₂ | 2.28 | SnR ₄ | 2.17 | 4 |
| SnCl ₂ | 2.42~2.43 | SnCl ₄ | 2.28~2.31 | 5-8 |
| SnBr ₂ | 2.55 | SnBr ₄ | 2.44 | 5,6,9 |
| SnI ₂ | 2.73~2.78 | SnI ₄ | 2.64 | 5,6,9 |

Table 7.3. Ionic radii (Å)

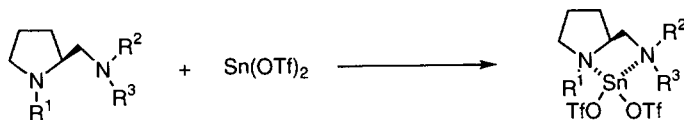
| Oxidation state | Ionic radii |
|-----------------|-------------|
| Sn(II) | 1.02 |
| Sn(IV) | 0.71 |

characteristics are utilized in the following asymmetric reactions: to form a rigid complex with chiral diamines and to create an efficient asymmetric field, which has enough space to include several reactants. This chapter surveys use of tin(II) and tin(IV) Lewis acids, especially chiral Lewis acid, in organic synthesis.

2. Asymmetric aldol reaction

2.1 Preparation of chiral tin(II) Lewis acids

Chiral tin(II) Lewis acids are prepared *in situ* by mixing tin(II) triflate and chiral diamines, which are readily prepared from (*S*)-prolin in an appropriate solvent (Scheme 7.1). Tin(II) has three vacant orbitals, and after co-ordination of two nitrogen atoms one vacant orbital still remains.¹⁴ Therefore, chiral diamine-co-ordinated tin(II) has a rigid bicyclo[3.3.0]octane-like structure consisting of two fused five-membered rings, and can activate an aldehyde using the vacant orbital without changing the rigid structure.¹⁵

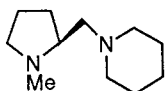
**Scheme 7.1**

2.2 Asymmetric aldol reactions of silyl enol ethers derived from acetic and propionic acid derivatives

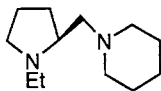
In the presence of a stoichiometric amount of tin(II) triflate, chiral diamine **1**, and tributyltin fluoride (Bu₃SnF), 1-*S*-ethyl-1-trimethylsiloxyethene or 1-*S*-*t*-butyl-1-trimethylsiloxyethene reacts with aldehydes to afford the corresponding adducts in high yields with high enantioselectivities.¹⁶ No chiral induction is observed without using Bu₃SnF. Although the precise function of Bu₃SnF is not yet clarified, it is believed that the fluoride connects the chiral tin(II) Lewis acid with the nucleophile, the silyl enol ether.^{17,18}

In the reactions with the propionate derivatives, which provide synthetically useful α-methyl-β-hydroxy ester derivatives, combination of tin(II) triflate, chiral diamine **5**, and Bu₂Sn(OAc)₂ gives better results.^{17,19} The asymmetric

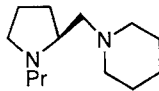
7: Tin Lewis acid



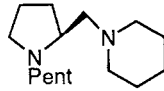
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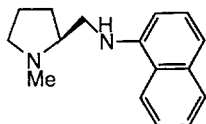
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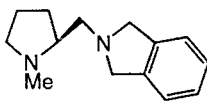
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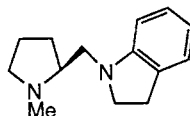
4



5



6



7

aldol reactions proceed with higher enantioselectivity, and in addition, the reactions proceed faster using $\text{Bu}_2\text{Sn}(\text{OAc})_2$ than using Bu_3SnF as an additive.²⁰ A wide variety of aldehydes including aliphatic, aromatic, and α,β -unsaturated aldehydes are applicable to this reaction, and in all cases the aldol adducts are obtained in high yields with perfect *syn*-selectivities, and the enantiomeric excesses of these *syn*-adducts are more than 98%.

Protocol 1.

Asymmetric aldol reaction of (Z)-1-S-ethyl-1-trimethylsiloxypropene (8Z) with an aldehyde

Equipment

- Two-necked, round-bottomed flask (30 mL) fitted with a magnetic stirring bar. A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Vacuum/inert gas source (argon source may be an argon balloon)

Materials

- Tin(II) triflate (FW 416.82), 166.3 mg, 0.40 mmol
- Distilled chiral diamine **5** (FW 240.35), 115.4 mg, 0.48 mmol
- Distilled $\text{Bu}_2\text{Sn}(\text{OAc})_2$ (FW 351.03), 154.5 mg, 0.44 mmol
- Distilled benzaldehyde (FW 106.12), 38.2 mg, 0.36 mmol
- Distilled (Z)-1-S-ethyl-1-trimethylsiloxypropene (**8Z**) (FW 190.38), 76.2 mg, 0.40 mmol
- Dry, distilled dichloromethane

1. Flame dry the reaction vessel with a stirring bar under dry argon.
2. Add tin(II) triflate to the flask in dry box under argon atmosphere and dry at 100 °C for 1 h under reduced pressure (0.5 mmHg).

Protocol 1. Continued

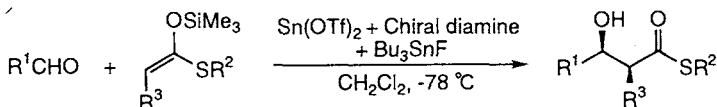
3. Add 1 ml of dry dichloromethane to the flask with stirring. Add chiral diamine **5** in 0.5 mL of dichloromethane and then $\text{Bu}_2\text{Sn}(\text{OAc})_2$ in 0.5 mL of dichloromethane.
 4. After stirring the mixture for 5 min at room temperature, cool to -78°C .
 5. Add (Z)-1-S-ethyl-1-trimethylsiloxypropene (**8Z**) in 0.5 mL of dichloromethane and benzaldehyde in 0.5 mL of dichloromethane successively.
 6. After stirring for 20 h, add saturated sodium hydrogen carbonate. Warm the mixture to the room temperature, filter the suspension through a pad of Celite, and wash the filter cake three times with dichloromethane. Extract the aqueous layer three times with dichloromethane and wash the combined extracts with brine.
 7. Concentrate the crude product under reduced pressure and apply to preparative TLC (silica gel). Elute with a mixed solvent of hexane–ethyl acetate (4:1) to obtain the product. Only *syn* isomer is produced and the optical purity is determined by HPLC analysis using a chiral column.
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

2.3 Asymmetric synthesis of 2-substituted malate derivatives

2-Substituted malic acids and their esters are widely distributed in nature as a variety of natural sources produced by plants or micro-organisms. Of more interest is their common inclusion as carboxylic acid components in biologically active alkaloids, and intense efforts have been made to prepare their carboxylic acid residues as optically active forms. Concerning asymmetric synthesis of 2-substituted malates, asymmetric aldol reactions of acetic acid derivatives with α -ketoesters are one of the most prospective methods.^{21–23}

In the presence of tin(II) triflate, (S)-1-pentyl-2-[(piperidin-1-yl)methyl]-pyrrolidine (**4**), and Bu_3SnF , 1-S-ethyl-1-trimethylsiloxyethene reacts with methyl pyruvate to give the desired adduct in 92% ee. Methyl isopropylglyoxylate and methyl phenylglyoxylate also react with 1-S-ethyl-1-trimethylsiloxyethene to give the corresponding 2-substituted malates in good yields with excellent enantioselectivities.²⁴

Recently, a series of pyrrolizidine alkaloids has attracted much attention due to their potent hepatotoxic, carcinogenic, and mutagenic properties.^{25–30} These alkaloids, especially 11- or 12-membered pyrrolizidine dilactones such as integerrimine, senecionine, fulvine, crispatine, etc., possess unique common structures of the α -methyl- β -hydroxy- β -alkyl units, and rather complicated multistage transformations have often been required for the stereoselective construction of these successive asymmetric centres including quaternary carbons.

Table 7.4. Asymmetric aldol reactions of silyl enolates with aldehydes

| Entry | R ¹ | R ² | R ³ | Chiral diamine | Yield(%) | syn/anti | ee (%) |
|-------|--|----------------|----------------|----------------|----------|----------|--------|
| 1 | Ph | Et | H | 1 | 78 | — | 82 |
| 2 | Ph | Et | H | 5 | 52 | — | 92 |
| 3 | Ph | t-Bu | H | 1 | 73 | — | 86 |
| 4 | Ph(CH ₂) ₂ | Et | H | 1 | 70 | — | 78 |
| 5 | Ph(CH ₂) ₂ | Et | H | 5 | 50 | — | 81 |
| 6 | Ph(CH ₂) ₂ | t-Bu | H | 1 | 71 | — | 85 |
| 7 | <i>i</i> -Pr | Et | H | 1 | 77 | — | 95 |
| 8 | <i>t</i> -Bu | Et | H | 1 | 90 | — | >98 |
| 9 | Ph | Et | Me | 5 | 85 | 100/0 | >98 |
| 10 | <i>p</i> -Cl Ph | Et | Me | 5 | 96 | 100/0 | >98 |
| 11 | <i>p</i> -Me ₃ Ph | Et | Me | 5 | 92 | 100/0 | >98 |
| 12 | <i>p</i> -MeO ₃ Ph | Et | Me | 5 | 95 | 100/0 | >98 |
| 13 | CH ₃ (CH ₂) ₆ | Et | Me | 5 | 90 | 100/0 | >98 |
| 14 | <i>c</i> -C ₆ H ₁₁ | Et | Me | 5 | 90 | 100/0 | >98 |
| 15 | <i>i</i> -Pr | Et | Me | 5 | 70 | 100/0 | >98 |
| 16 | <i>i</i> -Bu | Et | Me | 5 | 86 | 100/0 | >98 |
| 17 | (<i>E</i>)-CH ₃ CH=CH | Et | Me | 5 | 92 | 100/0 | >98 |
| 18 | (<i>E</i>)-PhCH=CH | Et | Me | 5 | 91 | 100/0 | >98 |
| 19 | (<i>E</i>)- <i>n</i> -PrCH=CH | Et | Me | 5 | 91 | 100/0 | >98 |
| 20 |  | Et | Me | 5 | 93 | 100/0 | >98 |
| 21 |  | Et | Me | 5 | 92 | 100/0 | >98 |

Protocol 2.**Asymmetric aldol reaction of 1-*S*-ethyl-1-trimethylsiloxyethene with an α-ketoester***Equipment*

- Two-necked, round bottomed flask (30 mL) fitted with a magnetic stirring bar. A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Vacuum/inert gas source (argon source may be an argon balloon)

Materials

- Tin(II) triflate (FW 416.82), 166.7 mg, 0.40 mmol
- Distilled chiral diamine **4** (FW 238.42), 114.4 mg, 0.48 mmol
- Bu₃SnF (FW 309.05), 136.0 mg, 0.44 mmol

Protocol 2. Continued

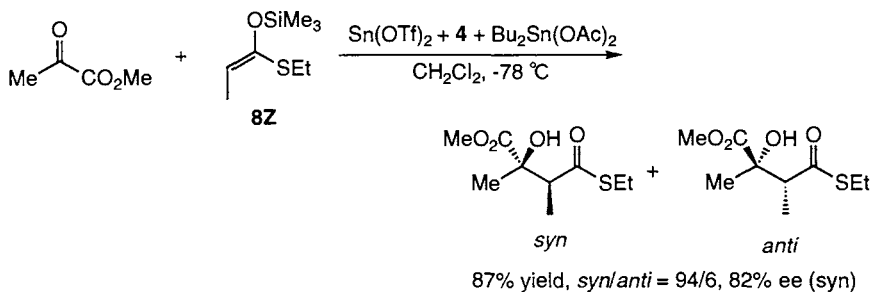
- Distilled methylpyruvate (FW 102.09), 36.8 mg, 0.36 mmol
- Distilled 1-*S*-ethyl-1-trimethylsiloxyethene (FW 176.35), 70.5 mg, 0.40 mmol
- Dry, distilled dichloromethane

1. Flame dry the reaction vessel with a stirring bar under dry argon.
2. Add tin(II) triflate to the flask in dry box under argon atmosphere and dry at 100°C for 1 h under reduced pressure (0.5 mmHg).
3. Add 1 ml of dry dichloromethane to the flask with stirring. Add chiral diamine **4** in 1 mL of dichloromethane and then Bu₃SnF.
4. After stirring the mixture for 5 min at room temperature, cool to -78°C.
5. Add 1-*S*-ethyl-1-trimethylsiloxyethene in 0.5 mL of dichloromethane and methylpyruvate in 0.5 mL of dichloromethane successively.
6. After stirring for 20 h, add saturated sodium hydrogen carbonate. Warm the mixture to the room temperature and filter the suspension through a pad of Celite, and wash the filter cake three times with dichloromethane. Extract the aqueous layer three times with dichloromethane and wash the combined extracts with brine.
7. Concentrate the crude product under reduced pressure and apply to preparative TLC (silica gel). Elute with a mixed solvent of hexane-ethyl acetate (4:1) to obtain the product. The optical purity is determined by HPLC analysis using a chiral column.

When (*Z*)-1-*S*-ethyl-1-trimethylsiloxypropene (**8Z**) is treated with methylpyruvate in the presence of tin(II) triflate, (*S*)-1-pentyl-2-[(piperidin-1-yl)methyl]pyrrolidine (**4**), and Bu₃SnF, the reaction proceeds smoothly to give the *syn* isomer in high yield with high diastereo- and enantioselectivities (Scheme 7.2).³¹ Similarly, (*Z*)-1-*S*-ethyl-1-trimethylsiloxypropene (**8Z**) reacts

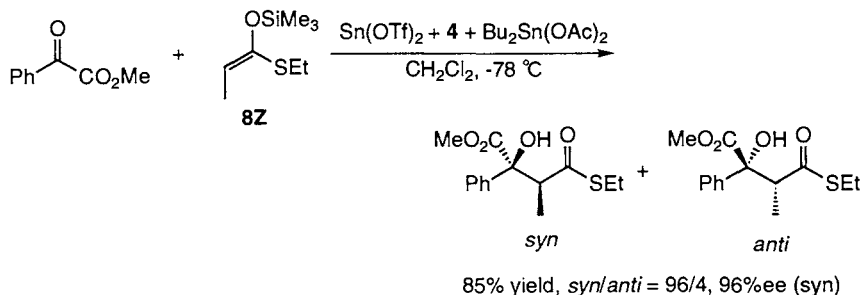
Table 7.5. Asymmetric aldol reactions of α-Ketoesters

| $ \begin{array}{c} \text{O} \\ \parallel \\ \text{R}-\text{C}-\text{CO}_2\text{Me} \end{array} + \begin{array}{c} \text{OSiMe}_3 \\ \parallel \\ \text{C}=\text{C}-\text{SEt} \end{array} \xrightarrow[\text{CH}_2\text{Cl}_2, -78^\circ\text{C}]{\text{Sn(OTf)}_2 + \text{Chiral diamine} + \text{Bu}_3\text{SnF}} \begin{array}{c} \text{MeO}_2\text{C} \quad \text{OH} \quad \text{O} \\ \quad \quad \quad \parallel \\ \text{R}-\text{C}-\text{CH}_2-\text{C}-\text{SEt} \end{array} $ | | | | |
|--|--------------|----------------|-----------|--------|
| Entry | R | Chiral diamine | Yield (%) | ee (%) |
| 1 | Me | 1 | 73 | 83 |
| 2 | Me | 4 | 78 | 92 |
| 3 | <i>i</i> -Pr | 1 | 76 | >98 |
| 4 | <i>i</i> -Pr | 4 | 81 | >98 |
| 5 | Ph | 1 | 74 | >98 |
| 6 | Ph | 4 | 74 | >98 |



Scheme 7.2

with methyl phenylglyoxylate smoothly to give the corresponding *syn* adduct in a high enantiomeric excess (Scheme 7.3). The successive asymmetric centres including quarternary carbons are constructed efficiently with high selectivities by using this methodology.³² On the other hand, (*E*)-1-*S*-ethyl-1-trimethylsiloxypropene reacts with methyl phenylglyoxylate or methyl pyruvate very slowly under the same reaction conditions.

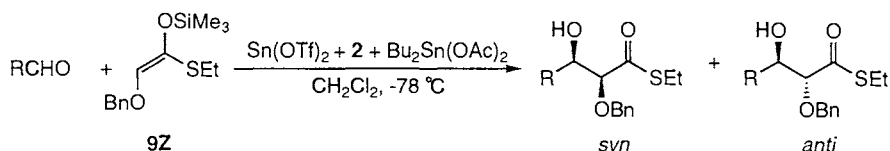


Scheme 7.3

2.4 Asymmetric synthesis of *syn*- and *anti*-1,2diol derivatives

Optically active 1,2-diol units are often observed in nature such as carbohydrates, macrolides, polyethers, etc. Recently several excellent asymmetric oxidation reactions of olefins using osmium tetroxide with a chiral ligand has been developed to achieve high enantiomeric excesses.^{33–36}

Asymmetric synthesis of 1,2-diol derivatives based on asymmetric aldol reactions of α -alkoxy silyl enol ethers with aldehydes has been developed. The reaction of (*Z*)-2-benzyloxy-1-*S*-ethyl-1-trimethylsiloxyethene (**9Z**) with benzaldehyde is carried out in dichloromethane at -78°C by using a chiral promoter consisting of tin(II) triflate, (*S*)-1-methyl-2-[(*N*-1-naphthylamino) methyl]-pyrrolidine (**5**), and Bu_3SnF , to afford the corresponding aldol adduct

Table 7.6. Asymmetric synthesis of *anti*-1,2-diol derivatives


| Entry | R | Yield (%) | syn/anti | ee (%) (<i>anti</i>) |
|-------|---|-----------|----------|------------------------|
| 1 | Ph | 83 | 1/99 | 96 |
| 2 | CH ₃ CH ₂ | 72 | 2/98 | 97 |
| 3 | c-C ₆ H ₁₁ | 59 | 9/91 | 96 |
| 4 | (<i>E</i>)-PhCH=CH | 88 | 2/98 | 98 |
| 5 | (<i>E</i>)-CH ₃ CH=CH | 85 | 2/98 | 97 |
| 6 | (<i>E,E</i>)-CH ₃ CH=CHCH=CH | 83 | 2/98 | 95 |

in a 69% yield with *anti* preference. The enantiomeric excesses of *syn* and *anti* aldols are 30% and 97%, respectively. Examination of several chiral diamines improves the diastereoselectivity, and when (*S*)-1-ethyl-2-[(piperidin-1-yl)-methyl]pyrrolidine (**2**) is employed, the aldol adduct is obtained in a 54% yield with excellent diastereo- and enantioselectivities. Furthermore, the yield is improved without any loss of the stereoselectivity by the combination of tin(II) triflate, **2**, and Bu₂Sn(OAc)₂.

The results employing several kinds of aldehydes such as aromatic, aliphatic, α,β-unsaturated aldehydes and a dienal of the present asymmetric aldol reaction are shown in Table 7.6.³⁷ In all cases, *anti*-α,β-dihydroxy thioesters are obtained in high yields with excellent diastereo- and enantioselectivities. The aldol adducts thus obtained, optically active *anti*-α,β-dihydroxy ester derivatives, are generally difficult to prepare by the conventional asymmetric oxidative procedure because starting materials, *cis*-α,β-unsaturated ester equivalents, are sometimes difficult to obtain. Moreover, a consideration of the mechanistic model of the asymmetric dioxyosmylation was recently reported and it was shown that preparation of *anti*-1,2-diols in high enantiomeric excesses is difficult.³⁸ According to the present aldol methodology, two hydroxyl groups can be introduced in 1,2-*trans* position stereoselectively during new carbon-carbon bond formation.

In contrast to the fact the aldol reactions of the silyl enol ether derived from *S*-ethyl propanethioate (**8Z**) with aldehydes using the above chiral promoter proceed with *syn* preference in excellent diastereo- and enantioselectivities (see Table 7.4), excellent *anti* selectivities have been achieved in reactions of (*Z*)-2-benzyloxy-1-*S*-ethyl-1-trimethylsiloxyethene (**9Z**) with aldehydes. The consideration of the transition states of these aldol reactions has led to the assumption that (i) co-ordination of the oxygen atom of the α-benzyloxy group of silyl enol ether **9Z** to the tin(II) atom of tin(II) triflate is essential in

Table 7.7. Asymmetric synthesis of *syn*-1,2-diol derivatives

| $\text{RCHO} + \begin{array}{c} \text{OSiMe}_3 \\ \\ \text{C} = \text{C} \\ \quad \\ \text{TBSO} \quad \text{SEt} \end{array} \xrightarrow[\text{CH}_2\text{Cl}_2, -78^\circ\text{C}]{\text{Sn}(\text{OTf})_2 + \mathbf{3} + \text{Bu}_2\text{Sn}(\text{OAc})_2} \begin{array}{c} \text{HO} \quad \text{O} \\ \quad \\ \text{R} \quad \text{C} \\ \quad \\ \text{OTBS} \quad \text{SEt} \end{array} + \begin{array}{c} \text{HO} \quad \text{O} \\ \quad \\ \text{R} \quad \text{C} \\ \quad \\ \text{OTBS} \quad \text{SEt} \end{array}$ | | | | |
|---|---|-----------|-----------------|------------------------|
| Entry | R | Yield (%) | <i>syn/anti</i> | ee (%) (<i>anti</i>) |
| 1 | Ph | 86 | 88/12 | 90 |
| 2 | CH ₃ CH ₂ | 46 | 92/8 | 82 |
| 3 |  | 93 | 94/6 | 93 |
| 4 | (<i>E</i>)-PhCH=CH | 76 | 90/10 | 92 |
| 5 | (<i>E</i>)-CH ₃ CH=CH | 75 | 97/3 | 94 |
| 6 | (<i>E,E</i>)-CH ₃ CH=CHCH=CH | 83 | 93/7 | 94 |

the *anti* selective transition state, leading to the different course in the diastereofacial selectivity compared with that of the *syn* selective reaction of **8Z**, (ii) *syn* α,β -dihydroxy thioesters would be formed when this co-ordination is restrained.

According to this hypothesis, the *t*-butyldimethylsilyl group has been chosen as a sterically hindered functional group, which would forbid the co-ordination of the oxygen atom to tin(II) atom, and (*Z*)-2-*t*-butyldimethylsiloxy-1-*S*-ethyl-1-trimethylsiloxyethene (**10Z**) has been prepared. As expected, the *syn* aldol adduct has been obtained under the same reaction conditions. In the presence of tin(II) triflate, chiral diamine (*S*)-1-propyl-2-[(piperidin-1-yl)methyl]pyrrolidine (**3**), and Bu₂Sn(OAc)₂, the reaction of **9Z** with benzaldehyde proceeds smoothly to give the corresponding aldol adduct in high yield with high *syn* selectivity.

Several examples of the *syn* selective aldol reactions are shown in Table 7.7. In all cases, the reactions proceed smoothly to afford the aldol adducts in good yields with very high *syn* selectivities, and the enantiomeric excesses of these *syn* isomers are more than 90% in most cases.^{31,39}

Now it becomes possible to control the enantiofacial selectivity of the silyl enol ethers derived from α -alkoxy thioesters **9Z** and **10Z** by choosing the appropriate protective groups of the alkoxy parts of the silyl enol ethers, and both diastereomers of optically active α,β -dihydroxy thioesters can be synthesized.

It is also possible to synthesize both enantiomers including 1,2-diol units with perfect stereochemical control by using similar types of chiral sources, based on Chiral Lewis Acid-Controlled Synthesis (CLAC Synthesis).^{40–43} The key is to use two new types of chiral diamines, **6** and **7**. In the presence of tin(II) triflate, chiral diamine **6**, and Bu₂Sn(OAc)₂, (*Z*)-2-(*t*-butyldimethylsiloxy)-1-ethylthio-1-trimethylsiloxyethene (**10Z**) reacts with aldehydes to

afford the desired aldol adducts with a *2S*, *3R* configuration. On the other hand, when a similar chiral diamine, **7**, is used, the reaction also proceeds smoothly, but the absolute configuration of the adducts is the reverse (*2R*, *3S*). In both cases, the selectivities are very high; almost perfect *syn*-selectivities and more than 98% enantiomeric excesses are obtained. Diamines **6** and **7** are readily prepared from *S*-proline and the absolute configurations of the 2-position were *S* in both cases. The difference is the position of the benzene ring connected to the pyrrolidine moiety. It is exciting that the slight difference in the structure of the chiral sources completely reverses the enantiofacial selectivities.

Protocol 3.

Synthesis of both enantiomers of 1,2-diol units by choosing similar chiral ligands in asymmetric aldol reactions of the silyl enol ethers with aldehydes (chiral Lewis acid-controlled synthesis)

Equipment

- Two-necked, round-bottomed flask (30 mL) fitted with a magnetic stirring bar. A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Vacuum/inert gas source (argon source may be an argon balloon)

Materials

- Tin(II) triflate (FW 416.82), 166.7 mg, 0.40 mmol
- Distilled chiral diamine **6** (FW 216.33), 103.8 mg, 0.48 mmol
- Distilled chiral diamine **7** (FW 216.33), 103.8 mg, 0.48 mmol
- Distilled Bu₂Sn(OAc)₂ (FW 351.03), 154.5 mg, 0.44 mmol
- Distilled benzaldehyde (FW 106.12), 38.2 mg, 0.36 mmol
- Distilled (*Z*)-2-(*t*-butyldimethylsiloxy)-1-ethylthio-1-trimethylsiloxyethene (**10Z**) (FW 278.53), 111.4 mg, 0.40 mmol
- Dry, distilled dichloromethane

1. Flame dry the reaction vessel with a stirring bar under dry argon.
2. Add tin(II) triflate to the flask in dry box under argon atmosphere and dry at 100°C for 1 h under reduced pressure (0.5 mmHg).
3. Add 1 mL of dry dichloromethane to the flask with stirring. Add chiral diamine **6** or **7** in 0.5 mL of dichloromethane and then Bu₂Sn(OAc)₂ in 0.5 mL of dichloromethane.
4. After stirring the mixture for 5 min at room temperature, cool to –78°C.
5. Add (*Z*)-2-(*t*-butyldimethylsiloxy)-1-ethylthio-1-trimethylsiloxyethene (**10Z**) in 0.5 mL of dichloromethane and benzaldehyde in 0.5 mL of dichloromethane successively.
6. After stirring for 20 h, add saturated sodium hydrogencarbonate. Warm the mixture to the room temperature and filter the suspension through a pad of

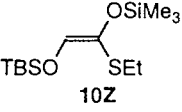
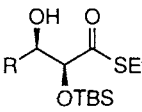
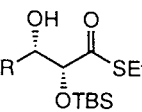



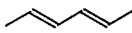



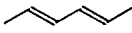
Celite, and wash the filter cake three times with dichloromethane. Extract the aqueous layer three times with dichloromethane and wash the combined extracts with brine.

- Concentrate the crude product under reduced pressure and apply to preparative TLC (silica gel). Elute with a mixed solvent of hexane–ethyl acetate (4:1) to obtain the product. Only *syn* isomer is produced and the optical purity is determined by HPLC analysis using a chiral column.

2.5 Catalytic asymmetric aldol reactions

As shown in the previous sections, highly diastereo- and enantioselective aldol reactions of silyl enol ethers with aldehydes using the novel promoter

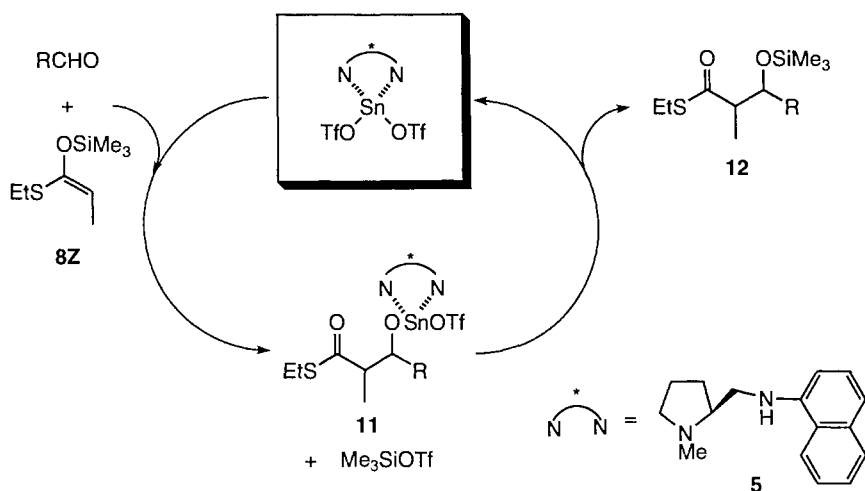
Table 7.8. Synthesis of both enantiomers

| $\text{RCHO} + \text{10Z} \xrightarrow[\text{Bu}_2\text{Sn}(\text{OAc})_2, \text{CH}_2\text{Cl}_2, -78^\circ\text{C}]{\text{Sn}(\text{OTf})_2 + \text{Chiral diamine}}$ <div style="text-align: center;">  <p>10Z</p> </div> | | | | | |
|--|----------------|-----------|--|-------------------------|---------------------|
| <div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;">  <p>(2<i>S</i>,3<i>R</i>)</p> </div> <div>+</div> <div style="text-align: center;">  <p>(2<i>R</i>,3<i>S</i>)</p> </div> </div> | | | | | |
| R | Chiral diamine | Yield (%) | 2 <i>S</i> ,3 <i>R</i> /2 <i>R</i> ,3 <i>S</i> | <i>syn/anti</i> | ee (%) ^a |
| Ph | 6 | 86 | 98/2 | 99.0/1.0 | (98) |
| C ₂ H ₅ | 6 | 61 | >99/1 | 99.0/1.0 | (98) |
|  | 6 | 86 | >99/1 | 99.0/1.0 | (98) |
| Ph  | 6 | 84 | >99/1 | 99.5/0.5 | (99) |
|  | 6 | 94 | >99/1 | 98.5/1.5 ^b | (97) |
|  | 6 | 86 | 98/2 | 98.0/2.0 | (96) |
| Ph | 7 | 82 | 99/1 | 1.0/99.0 | (98) |
| C ₂ H ₅ | 7 | 63 | >99/1 | 1.0/99.0 | (98) |
|  | 7 | 81 | >99/1 | 1.0/99.0 | (98) |
| Ph  | 7 | 80 | 99/1 | 1.0/99.0 | (98) |
|  | 7 | 86 | >99/1 | <0.5/>99.5 ^b | (>99) |
|  | 7 | 78 | >99/1 | <0.5/>99.5 | (>99) |

^a Enantiomeric excesses of *syn*-adducts.

^b 2*S*,3*S*/2*R*,3*R*.

system, combined use of stoichiometric amounts of tin(II) triflate, a chiral diamine, and a tin(IV) compound (Bu_3SnF or $\text{Bu}_2\text{Sn}(\text{OAc})_2$), have been developed. According to these reactions, optically active aldol adducts are easily prepared from both achiral aldehydes and silyl enol ethers, but stoichiometric use of the chiral source still remains a problem in terms of practical use. Based on the investigations to characterize the above promoter system as well as to clarify the mechanism of these reactions toward a truly catalytic aldol process,⁴⁴ the following catalytic cycle is postulated (Scheme 7.4).



Scheme 7.4. The catalytic cycle of the asymmetric aldol reaction

A chiral diamine co-ordinated tin(II) triflate (chiral tin(II) Lewis acid) interacts with an aldehyde, and tin(II) alkoxide **11** and trimethylsilyl triflate (TMSOTf) are initially produced by the attack of silyl enol ether **8Z** to the activated aldehyde. When the metal exchange between tin(II) and silicon of the above product **11** takes place smoothly, the corresponding aldol adduct can be obtained as its trimethylsilyl ether **12** together with the regeneration of the catalyst. If the above-mentioned metal exchange step is slow, undesirable TMSOTf-promoted reaction^{45,46} (to afford the achiral aldol adduct) proceeds to result in lowering the selectivity.

On the basis of these considerations, a slow addition of the substrates to the solution of the catalyst has been performed to keep the concentration of TMSOTf as low as possible during the reaction. A dichloromethane solution of silyl enol ether **8Z** and an aldehyde is added for over 9 h to a dichloromethane solution of the catalyst (20 mol%) (Table 7.9).⁴⁷ As expected, aldol adducts are obtained in good yields with excellent enantiomeric excesses and high diastereoselectivities in some cases, but the selectivities are not so high

Table 7.9. Catalytic asymmetric aldol reactions (1) (solvent : CH₂Cl₂)

| Entry | R | Yield (%) | syn/anti | ee (%) (anti) |
|-------|--|-----------|----------|---------------|
| 1 | Ph | 86 | 93/7 | 91 |
| 2 | <i>p</i> -Cl Ph | 80 | 93/7 | 93 |
| 3 | <i>p</i> -CH ₃ Ph | 82 | 78/22 | 80 |
| 4 | CH ₃ (CH ₂) ₆ | 75 | 100/0 | >98 |
| 5 | <i>c</i> -C ₅ H ₁₁ | 31 | >99/1 | 74 |
| 6 | (<i>E</i>)-CH ₃ CH=CH | 51 | 84/16 | 77 |
| 7 | (<i>E</i>)-CH ₃ (CH ₂) ₂ CH=CH | 62 | 81/19 | 74 |

when *p*-tolualdehyde, α,β -unsaturated aldehydes, and cyclohexanecarboxaldehyde are used.

The lower selectivities are considered to be ascribed to the incompleteness of the above catalytic cycle, especially the metal exchange reaction of the initially formed aldol adduct **11** with TMSOTf (metal exchange between tin(II) and silicon). In order to accelerate this metal exchange step, various polar solvents with low melting points (below -78°C) have been carefully examined by taking the reaction of **8Z** with benzaldehyde as a model, and finally propionitrile (C₂H₅CN) has been found to be an excellent solvent.⁴⁸ The examination of the addition time (addition of the reactants to the solution of the catalyst) reveals that the rate of the metal exchange in propionitrile is faster than that in dichloromethane. Although 9 h (addition time) is necessary to attain the best result in dichloromethane, comparable selectivities are achieved when the substrates are added to the catalyst for 3 h in propionitrile. It is noted that a tin(II) triflate is more soluble in propionitrile than in dichloromethane indicating that the co-ordination of the nitrile group to tin(II) is expected to be rather strong, but that the ligand exchange of the nitrile for the diamine takes place smoothly to form the desired chiral Lewis acid when the chiral diamine is added to this propionitrile solution of tin(II) triflate. Although tin(II) triflate is also soluble in tetrahydrofuran (THF) or 1,2-dimethoxyethane (DME), the above aldol reaction does not proceed at -78°C after the addition of chiral diamine **5** in these solvents.

Several aldehydes including aromatic, aliphatic, and α,β -unsaturated aldehydes, are applicable to this reaction, and the desired products are obtained in good yields with high selectivities (>90% ee, Table 7.10). In particular, the lower yields or selectivities observed in the reaction of *p*-tolualdehyde, (*E*)-crotonaldehyde, (*E*)-2-hexenal, and cyclohexanecarboxaldehyde, are remarkably improved by using propionitrile as a solvent. High selectivities are also achieved even when 10 mol% of the catalyst was used.⁴⁹

The key step in the asymmetric aldol reactions using the chiral tin(II) Lewis acid is believed to be the metal exchange process from tin(II) to silicon on the initially produced aldol adduct **11**. If this step is slow, the TMSOTf-promoted aldol reaction of aldehydes with silyl enol ethers proceed to result in lower

Table 7.10. Catalytic asymmetric aldol reactions (2) (solvent : C₂H₅CN)

| Entry | R | Addition time (h) | Yield (%) | syn/anti | ee (%) (syn) |
|-------|--|-------------------|-----------|----------|--------------|
| 1 | Ph | 3 | 77 | 92/8 | 90 |
| 2 | <i>p</i> -Cl Ph | 4.5 | 83 | 87/13 | 90 |
| 3 | <i>p</i> -CH ₃ Ph | 3 | 75 | 89/11 | 91 |
| 4 | CH ₃ (CH ₂) ₆ | 4.5 | 80 | 100/0 | >98 |
| 5 | <i>c</i> -C ₈ H ₁₁ | 3 | 71 | 100.0 | >98 |
| 6 | (<i>E</i>)-CH ₃ CH=CH | 3 | 76 | 96/4 | 93 |
| 7 | (<i>E</i>)-CH ₃ (CH ₂) ₂ CH=CH | 3 | 73 | 97/3 | 93 |

selectivities. Slow addition of the substrates to the catalyst in propionitrile has been successfully performed to suppress the undesirable reaction. It is also expected that higher stereoselectivities would be obtained when the Lewis acidity of TMSOTf is reduced by using an additive. This idea has been obtained from the recent report on the catalytic asymmetric aldol reaction using TMSOTf, a chiral diamine, and tin(II) oxide.⁵⁰ In this reaction, the lone-pair electrons of tin(II) oxide interact with TMSOTf to weaken the Lewis acidity of TMSOTf, and as a result high selectivities have been obtained. It would be possible to reduce the Lewis acidity of various triflates by using this interaction. On the basis of this idea, it has been found that in the presence of a novel chiral catalyst system consisting of tin(II) triflate, a chiral diamine, and tin(II) oxide, highly enantioselective aldol reactions of the silyl enol ether of *S*-ethyl ethanethioate or *S*-ethyl propanethioate with aldehydes proceed smoothly to afford the aldol adducts in high yields.

Protocol 4.

Catalytic asymmetric aldol reactions of the silyl enol ether derived from (*Z*)-1-*S*-ethyl-1-trimethylsiloxypropene with an aldehyde

Equipment

- Two-necked, round-bottomed flask (30 mL) fitted with a magnetic stirring bar. A three-way stopcock is fitted to the top of the flask and connected to a vacuum/argon source
- Vacuum/inert gas source (argon source may be an argon balloon)

Materials

- Tin(II) triflate (FW 416.82), 33.4 mg, 0.08 mmol
- Tin(II) oxide (FW 134.71), 21.6 mg, 0.16 mmol
- Distilled chiral diamine **5** (FW 240.35), 23.1 mg, 0.096 mmol
- Distilled benzaldehyde (FW 106.12), 42.5 mg, 0.40 mmol
- Distilled (*Z*)-1-*S*-ethyl-1-trimethylsiloxypropene (FW 190.38), 76.2 mg, 0.40 mmol
- Dry, distilled propionitrile

1. Flame dry the reaction vessel with a stirring bar under dry argon.

2. Add tin(II) triflate and tin(II) oxide to the flask in dry box under argon atmosphere and dry at 100°C for 1 h under reduced pressure (0.5 mmHg).
3. Add 1 mL of dry propionitrile to the flask with stirring. Add chiral diamine **4** in 0.5 mL of propionitrile.
4. After stirring the mixture for 5 min at room temperature, cool to -78°C.
5. Slowly add a mixture of (Z)-1-S-ethyl-1-trimethylsiloxypropene in 0.5 mL of propionitrile and benzaldehyde in 0.5 mL of propionitrile over 4 h (syringe pump).
6. After stirring for 1 h (after addition), add saturated sodium hydrogen carbonate. Warm the mixture to the room temperature and filter the suspension through a pad of Celite, and wash the filter cake three times with dichloromethane. Extract the aqueous layer three times with dichloromethane and wash the combined extracts with brine. Dry the combined organic layer over (Na₂SO₄).
7. Concentrate the crude product under reduced pressure and apply to preparative TLC (silica gel). Elute with a mixed solvent of hexane-ethyl acetate (6:1) to obtain the product. Only *syn* isomer is produced and the optical purity is determined by HPLC analysis using a chiral column. After a usual work-up, the aldol-type adduct was isolated as the corresponding trimethylsilyl ether. The enantiomeric excess of this aldol adduct was determined after converting it into the corresponding alcohol (tetrahydrofuran-1M HCl (20:1) at 0°C) to be 94% by HPLC analysis using a chiral stationary phase column.

3. Asymmetric cyanation reaction

Asymmetric cyanation reaction of aldehydes is an important process in organic synthesis and recently some excellent works have been made in this

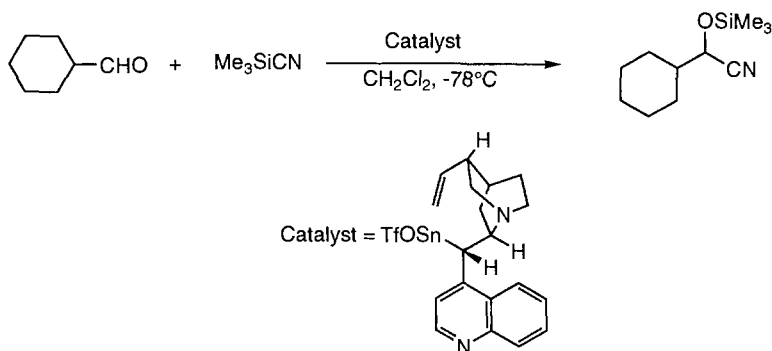
Table 7.11. Catalytic asymmetric aldol reactions using a novel catalyst system

| $R^1CHO + \begin{array}{c} \text{OSiMe}_3 \\ \diagup \\ R^2 \\ \diagdown \\ \text{SEt} \end{array}$ | | | $\xrightarrow[\text{C}_2\text{H}_5\text{CN, } -78^\circ\text{C, slow addition 4-6 h}]{20 \text{ mol\% Sn(OTf)}_2 + 5 + \text{SnO (20-40 mol\%)}}$ | | | $\begin{array}{c} \text{Me}_3\text{SiO} \quad \text{O} \\ \quad \quad \\ R^1 - \text{C} - \text{C} - \text{C} - \text{SEt} \\ \\ R^2 \\ \text{syn} \end{array}$ | | |
|---|---|----------------|---|-----------------|--------|--|-----------------|--------|
| Entry | R ¹ | R ² | With SnO | | | Without SnO | | |
| | | | Yield (%) | <i>syn/anti</i> | ee (%) | Yield (%) | <i>syn/anti</i> | ee (%) |
| 1 | Ph | Me | 78 | 95/5 | 93 | 77 | 93/7 | 90 |
| 2 | CH ₃ (CH ₂) ₆ | Me | 81 | 100/0 | >98 | 80 | 100/0 | >98 |
| 3 | CH ₃ CH=CH | Me | 85 | 99/1 | 95 | 76 | 96/4 | 93 |
| 4 | CH ₃ (CH ₂) ₂ CH=CH | Me | 81 | 100/0 | 94 | 73 | 97/3 | 93 |
| 5 | CH ₃ (CH ₂) ₃ | H | 71 | — | 92 | 79 | — | 91 |
| 6 | <i>i</i> -Pr | H | 50 | — | 92 | 48 | — | 90 |
| 7 | <i>c</i> -C ₆ H ₁₁ | H | 78 | — | 94 | 81 | — | 92 |
| 8 | CH ₃ (CH ₂) ₂ CH=CH | H | 83 | — | 84 | 65 | — | 72 |

area,^{51–55} however, some significant problems still remain such as stoichiometric use of chiral sources, lower enantioselectivities in the reaction with aliphatic aldehydes, etc.

For the synthesis of cyanohydrin trimethylsilyl ethers,^{56,57} one of the most convenient preparative methods is addition reactions of trimethylsilyl cyanide (TMS-CN) with aldehydes under the influence of Lewis acids such as zinc iodide (ZnI₂), aluminium chloride (AlCl₃), or under almost neutral conditions using anionic catalysis.^{58,59} It has been found that in the presence of a catalytic amount of Lewis base such as amine, phosphine, arsine, or antimony, TMS-CN smoothly reacts with aldehydes to afford the corresponding cyanohydrin trimethylsilyl ethers in excellent yields.

Asymmetric version of this reaction is investigated. In the presence of a catalytic amount of (+)-cinchonine (10 mol%), cyclohexanecarboxaldehyde reacts with TMS-CN smoothly at -78°C to give the corresponding cyanohydrin trimethylsilyl ether in a 94% yield with 25% ee. (+)-Cinchonine trimethylsilyl ether also works with the same level of ee (25%). The importance of the silyl ether is revealed as almost no chiral induction is observed when (+)-cinchonine acetate or benzoate is used as a catalyst. The proposed mechanism is that the tertiary amine part of (+)-cinchonine interacts with the silicon atoms of TMS-CN to form the pentavalent silicate and, at the same time, the Lewis acidic silicon atom of the silyl ether activates an aldehyde. This ‘double activation’ would provide a desirable transition state in the chiral induction using chiral catalysts. Based on this hypothesis, high levels of enantiomeric excesses are obtained when tin(II) ether is introduced instead of the silyl ether and a novel tin(II) Lewis acid, tin(II) monoalchoxymonotriflate containing (+)-cinchonine as a chiral source, has been designed. This chiral tin(II) Lewis acid is easily prepared from 1,1'-dimethylstannocene,⁶⁰ triflic acid, and (+)-cinchonine. In the presence of this Lewis acid, the reactions of TMS-CN with aldehydes proceed smoothly at -78°C in dichloromethane to give the corresponding cyanohydrin trimethylsilyl ether in high yields with good to excellent ees (Scheme 7.5). In the present reaction, the products are



Scheme 7.5

isolated as trimethylsilyl ether form and the reaction smoothly proceeds in the presence of 30 mol% of the tin(II) Lewis acid. The catalyst, tin(II) monoalkoxymonotriflate, is supposed to be regenerated from the initially produced tin(II) alkoxide and trimethylsilyl triflate.

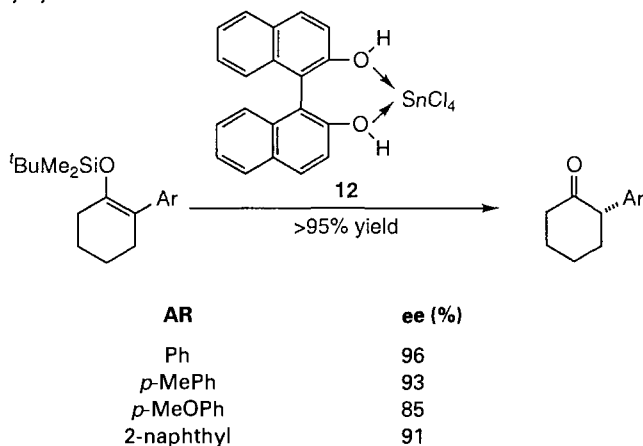
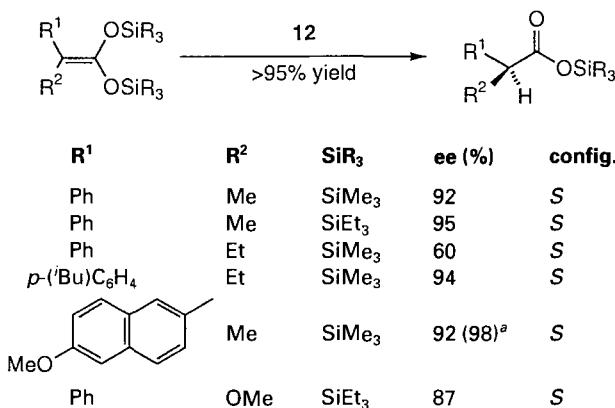
4. Asymmetric protonation

Enantioselective protonation of prochiral enol derivatives provides a convenient route for the preparation of optically active carbonyl compounds.^{61–89} Several examples of protonation of metal enolates by chiral proton sources^{64–79} and hydrolysis of enol esters catalysed by enzymes^{80,81} or antibodies⁸² under basic or neutral conditions, have been reported. The acid-promoted hydrolysis of enol ethers is an interesting alternative which has been little investigated for enantioselectivity.^{83–86} A new Lewis acid-assisted chiral Brønsted acid (LBA)^{87–89} for enantioselective protonation of prochiral silyl enol esters and ketene bis(trialkylsilyl) acetals is reported.⁹⁰

The LBA (**12**) is generated *in situ* from (*R*)-(+)-1,1'-binaphthol (BINOL) and tin tetrachloride (SnCl₄) in toluene or dichloromethane at -78°C . The protons in (*R*)-BINOL are activated by the co-ordination of tin tetrachloride. In the presence of a stoichiometric amount of **12** in dichloromethane, the C-protonation of the trimethylsilyl enol ether derived from 2-phenylcyclohexanone proceeds even at -78°C to form (*S*)-2-phenylcyclohexanone with good enantioselectivity (79% ee). The enantioselectivity of the reaction is dramatically increased by using sterically bulky *O*-substituents. The protonation of the *tert*-butyldimethylsilyl enol ether (Ar = Ph, SiR₃ = Si^tBuMe₂) occurs with excellent enantioselectivity (93% ee), and the best result (96% ee) is achieved by using toluene as solvent. (*R*)-BINOL is converted into a mixture of 2,2'-disiloxy-1,1'-binaphthyl and 2-siloxy-2'-hydroxy-1,1'-binaphthyl in dichloromethane solution, whereas it is only converted into the latter product in toluene. It is noted that the enantioselective protonation proceeded with a catalytic amount (10 mol %) of tin tetrachloride.

The enantioselective protonation of a variety of silyl enol ethers derived from 2-arylcyclohexanones with **12** under optimum conditions is summarized in Table 7.12. The reactions are generally complete after 1 h at -78°C . Excellent enantioselectivity was achieved in the reactions of the silyl enol ethers of 2-arylcyclohexanones, and both enantiomers can be obtained from racemic 2-arylcyclohexanones depending on the choice of optically active BINOL: (*S*)- and (*R*)-arylcyclohexanones are obtained using (*R*)- and (*S*)-**12**, respectively. Enantioselectivity is low in the case of the protonation of the silyl enol ether derived from 2-methylcyclohexanone (42% ee).

Representative results of application of **12** to enantioselective protonation of ketene bis(trialkylsilyl) acetals derived from 2-arylcarboxylic acids are summarized in Table 7.13. The crude carboxylic acids, which are formed with excellent enantiomeric excesses, are isolated as the corresponding methyl

Table 7.12. The enantioselective protonation of silyl enol ethers from 2-arylcyclohexanones with **12****Table 7.13.** Enantioselective protonation of ketene bis(trialkyl silyl) acetals derived from 2-arylcarboxylic acids^a After recrystallization from dichloromethane-hexane.

esters in quantitative yields. The enantioselectivity is independent of the steric features of trialkylsilyl substituents. Simple recrystallization of the (*S*)-methyl ester of naproxen can be used to upgrade the optical purity. It is noteworthy that the protonation of the ketene silyl acetal, diastereomer ratio (*E* and *Z*) = 63:37, derived from methyl 6-methoxy- α -methyl-2-naphthalene acetate and trimethylsilyl chloride by (*R*)-**12** gives the (*S*)-methyl ester of naproxen with decreased enantioselectivity (79% ee).

References

1. Poller, R. C. Organic compounds of group IV metals. In *Comprehensive organic chemistry*, Vol. 3 (ed. D. N. Jones). Pergamon Press, Oxford, **1979**, p 1073.
2. Davies, A. G.; Smith, P. J. Tin. In *Comprehensive organometallic chemistry*, Vol. 2 (ed. G. Wilkinson). Pergamon Press, Oxford, **1982**, p 519.
3. Pereyre, M.; Quintard, J.-P.; Rahm, A. *Tin in organic synthesis*, Butterworths, London, **1987**.
4. Mukaiyama, T. *Pure Appl. Chem.* **1986**, 58, 505.
5. Huheey, J. E. *Inorganic chemistry*, 3rd edn. Harper & Row, New York, **1983**.
6. Veith, M.; Recktenwald, O. *Top. Curr. Chem.* **1982**, 104, 1.
7. Goldberg, D. E. *J. Chem. Soc., Chem. Commun.* **1976**, 261.
8. Lister, M. W.; Sutton, L. E. *Trans. Far. Soc.* **1941**, 37, 406.
9. Akshin, P. A.; Spiridonov, V. P.; Khodshenko, A. N. *Zh. Fiz. Khim.* **1958**, 32, 1679.
10. Fuji, H.; Kimura, M. *Bull. Chem. Soc. Jpn* **1970**, 43, 1933.
11. Livingstone, R. L.; Rao, C. N. R. *J. Chem. Phys.* **1959**, 30, 399.
12. Lister, M. W.; Sutton, L. E. *Trans. Far. Soc.* **1941**, 37, 393.
13. Dean, J. A. *Lange's handbook of chemistry*, 3rd edn, McGraw-Hill, New York, **1985**, pp 3–121.
14. Donaldson, J. D. *Progress in Inorganic Chemistry*, (ed. F. A. Cotton, Vol. 8. John Wiley, New York, **1967**, p 287.
15. Shields, K. G.; Seccombe, R. C.; Kennard, C. H. L. *J. Chem. Soc., Dalton* **1973**, 741.
16. Kobayashi, S.; Mukaiyama, T. *Chem. Lett.* **1989**, 297.
17. Kobayashi, S.; Uchiro, H.; Fujishita, Y.; Shiina, I.; Mukaiyama, T. *J. Am. Chem. Soc.* **1991**, 113, 4247.
18. Mukaiyama, T.; Kobayashi, S. *J. Organomet. Chem.*, **1990**, 382, 39.
19. Mukaiyama, T.; Uchiro, H.; Kobayashi, S. *Chem. Lett.* **1989**, 1757.
20. Mukaiyama, T.; Uchiro, H.; Kobayashi, S. *Chem. Lett.* **1989**, 1001.
21. Ojima, I.; Yoshida, K.; and Inaba, S. *Chem. Lett.* **1977**, 429.
22. Mioskowski, C.; Solladie, G. *Tetrahedron* **1980**, 36, 227.
23. Stevens, R. W.; Mukaiyama, T. *Chem. Lett.* **1983**, 1799.
24. Kobayashi, S.; Fujishita, Y.; Mukaiyama, T. *Chem. Lett.* **1989**, 2069.
25. Robins, D. J. *Fortschr. Chem. Org. Naturst.* **1982**, 41, 115.
26. Mulzer, J.; Kirstein, H. M.; Buschmann, J.; Lehmann, C.; Luger, P. *J. Am. Chem. Soc.* **1991**, 113, 910.
27. Hirota, M.; Sakurai, Y.; Horita, K.; Yonemitsu, O. *Tetrahedron Lett.* **1990**, 31, 6367.
28. White, J. D.; Jayasinghe, L. R. *Tetrahedron Lett.* **1988**, 29, 2139.
29. Niwa, H.; Miyachi, Y.; Uosaki, T.; Yamada, K. *Tetrahedron Lett.* **1986**, 27, 4601.
30. Narasaka, K.; Sakakura, T.; Uchimaru, T.; Guedin-Vuong, D. *J. Am. Chem. Soc.* **1984**, 106, 2954.
31. Kobayashi, S.; Horibe, M.; Saito, Y. *Tetrahedron* **1994**, 50, 9629.
32. Kobayashi, S.; Hachiya, I. *J. Org. Chem.* **1992**, 57, 1324.
33. Tomioka, K.; Nakajima, M.; Koga, K. *J. Am. Chem. Soc.* **1987**, 109, 6213.
34. Jacobsen, E. N.; Marco, I.; Mungall, W. S.; Schroder, G.; Sharpless, K. B. *J. Am. Chem. Soc.* **1988**, 110, 1968.

35. Wai, J. S. M.; Marco, I.; Svendsen, S.; Finn, M. G.; Jacobsen, E. N.; Sharpless, K. B. *J. Am. Chem. Soc.* **1989**, *111*, 1123.
36. Johnson, R. A.; Sharpless, K. B. Catalytic asymmetric dihydroxylation. In *Catalytic asymmetric synthesis* (ed. I. Ojima). VCH, Weinheim, Germany, **1993**, pp 227–272.
37. Mukaiyama, T.; Uchiro, H.; Shiina, I.; Kobayashi, S. *Chem. Lett.* **1990**, 1019.
38. Corey, E. J.; Jardine, P. D.; Virgil, S.; Yuen, P.-W.; Connell, R. D. *J. Am. Chem. Soc.* **1989**, *111*, 9243.
39. Mukaiyama, T.; Shiina, I.; Kobayashi, S. *Chem. Lett.* **1991**, 1902.
40. Kobayashi, S.; Horibe, M. *J. Am. Chem. Soc.* **1994**, *116*, 9805.
41. Kobayashi, S.; Horibe, M. *Chem. Eur. J.* **1977**, *3*, 1472.
42. Kobayashi, S.; Hayashi, T. *J. Org. Chem.* **1995**, *60*, 1098.
43. Kobayashi, S.; Horibe, M.; Matsumura, M. *Synlett* **1995**, 675.
44. Kobayashi, S.; Harada, T.; Han, J. S. *Chem. Express* **1991**, *6*, 563.
45. Murata, S.; Suzuki, M.; Noyori, R. *Tetrahedron* **1988**, *44*, 4259.
46. Mukai, C.; Hashizume, S.; Nagami, K.; Hanaoka, M. *Chem. Pharm. Bull.* **1990**, *38*, 1509.
47. Mukaiyama, T.; Kobayashi, S.; Uchiro, H.; Shiina, I. *Chem. Lett.* **1990**, 129.
48. Kobayashi, S.; Fujishita, Y.; Mukaiyama, T. *Chem. Lett.* **1990**, 1455.
49. Kobayashi, S.; Uchiro, H.; Shiina, I.; Mukaiyama, T. *Tetrahedron* **1993**, *49*, 1761.
50. Mukaiyama, T.; Uchiro, H.; Kobayashi, S. *Chem. Lett.* **1990**, 1147.
51. Minamikawa, H.; Hayakawa, S.; Yamada, T.; Iwasawa, N.; Narasaka, K. *Bull. Chem. Soc. Jpn.* **1988**, *61*, 4379.
52. Tanaka, K.; Mori, A.; Inoue, S. *J. Org. Chem.* **1990**, *55*, 181.
53. Almsick, A.; Buddrus, J.; H-Schmidt, P.; Laumen, K.; Schneider, M. P. *J. Chem. Soc. Commun.* **1989**, 1391.
54. Ziegler, T.; Horsch, B.; Ellenberger, F. *Synthesis* **1990**, 575.
55. Hayashi, M.; Matsuda, T.; Oguni, N. *J. Chem. Soc., Chem. Commun.* **1990**, 1364.
56. Weber, W. P. *Silicon reagents for organic synthesis*. Springer-Verlag, Berlin, **1983**.
57. Colvin, E. W. *Silicon in organic synthesis*, Butterworths, London, **1981**.
58. Evans, D. A.; Truesdale, L. K. *Tetrahedron Lett.*, **1973**, 4929.
59. Evans, D. A.; Truesdale, L. K.; Carroll, G. L. *J. Chem. Soc., Chem. Commun.* **1973**, 55.
60. Mukaiyama, T.; Ichikawa, J.; Asami, M. *Chem. Lett.* **1983**, 293.
61. Duhamel, L.; Duhamel, P.; Launay, J.; Plaquevent, J. C. *Bull. Soc. Chim. Fr.* **1984**, II-421.
62. Fehr, C. *Chimia* **1991**, *45*, 253.
63. Waldmann, H. *Nachr. Chem. Tech. Lab.* **1991**, *39*, 413.
64. Duhamel, L.; Fouquay, S.; Plaquevent, J. C. *Tetrahedron Lett.* **1986**, *27*, 4975.
65. Hogeveen, H.; Eleveld, M. B. *Tetrahedron Lett.* **1986**, *27*, 631.
66. Gerlach, U.; Hunig, S. *Angew. Chem., Int. Ed. Engl.* **1987**, *26*, 1283.
67. Potin, D.; Williams, K.; Rebek, J., Jr. *Angew. Chem., Int. Ed. Engl.* **1990**, *29*, 1420.
68. Kumar, A.; Salunkhe, R. V.; Rane, R. A.; Dike, S. Y. *J. Chem. Soc., Chem. Commun.* **1991**, 485.
69. Henin, F.; Muzart, J.; Pete, J.-p.; Piva, O. *New J. Chem.* **1991**, *15*, 611.
70. Matsumoto, K.; Ohat, H. *Tetrahedron Lett.* **1991**, *32*, 4729.
71. Henin, F.; Muzart, J. *Tetrahedron: Asymmetry* **1992**, *3*, 1161.
72. Takeuchi, S.; Miyoshi, N.; Hirata, K.; Hayashida, H. Ohgo, Y. *Bull. Chem. Soc. Jpn.* **1992**, *65*, 2001.

73. Yasukata, T.; Koga, K. *Tetrahedron: Asymmetry* **1993**, 4, 35.
74. Fuji, K.; Tanaka, K.; Miyamoto, H. *Tetrahedron: Asymmetry* **1993**, 4, 247.
75. Haubenreich, T.; Hunig, S.; Schultz, H-J. *Angew. Chem., Int. Ed. Engl.* **1993**, 32, 398.
76. Fehr, C.; Stempf, I.; Galindo, J. *Angew. Chem., Int. Ed. Engl.* **1993**, 32, 1044.
77. Vedejs, E.; Lee, N.; Sakata, S. T. *J. Am. Chem. Soc.* **1994**, 116, 2175.
78. Yanagisawa, A.; Kuribayashi, T.; Kikuchi, T.; Yamamoto, H. *Angew. Chem. Int. Ed. Engl.* **1994**, 33, 107.
79. Aboulhoda, S. J.; Henin, F.; Muzart, J.; Thorey, C.; Behnen, W.; Martens, J.; Mehler, T. *Tetrahedron: Asymmetry* **1994**, 5, 1321.
80. Matsumoto, K.; Tsutsumi, S.; Ihori, T.; Ohta, H. *J. Am. Chem. Soc.* **1990**, 112, 9614.
81. Kume, Y.; Ohta, H. *Tetrahedron Lett.* **1992**, 33, 6367.
82. Fuji, I.; Lerner, R. A.; Janda, K. D. *J. Am. Chem. Soc.* **1991**, 113, 8528.
83. Cavelier, F.; Gomez, R.; Jacquier, R.; Verducci, J. *Tetrahedron: Asymmetry* **1993**, 4, 2501.
84. Cavelier, F.; Gomez, R.; Jacquier, R.; Verducci, J. *Tetrahedron Lett.* **1994**, 35, 2891.
85. Reymond, J.-L.; Reber, J.-L.; Lerner, R. A. *Angew. Chem., Int. Ed. Engl.* **1994**, 33, 475.
86. Novice, M. H.; Seikaly, H. R.; Seiz, A. D.; Tidwell, T. T. *J. Am. Chem. Soc.* **1980**, 102, 5835.
87. Ishihara, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1994**, 116, 1561.
88. Ishihara, K.; Miyata, M.; Hattori, K.; Tada, T.; Yamamoto, H. *J. Am. Chem. Soc.* **1994**, 116, 10520.
89. Kobayashi, S.; Araki, M.; Hachiya, I. *J. Org. Chem.* **1994**, 59, 3758.
90. Ishihara, K.; Kaneeda, M.; Yamamoto, H. *J. Am. Chem. Soc.* **1994**, 116, 11179.

Silicon(IV) reagents

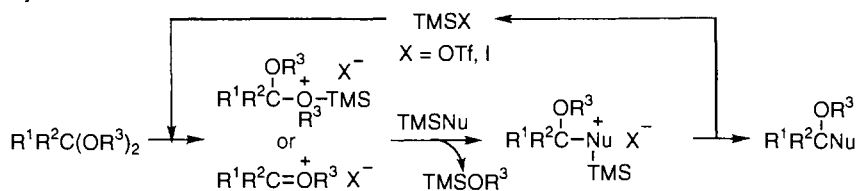
AKIRA HOSOMI and KATSUKIYO MIURA

1. Introduction

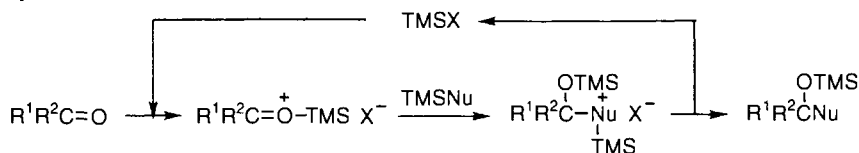
Organosilicon compounds are one of the most widely and frequently used organometallics in organic synthesis. They are easy to handle and synthesize due to their thermal and aerial stability, and exhibit a variety of reactivities derived from the electropositive nature of silicon as well as the steric and electronic effects of the silyl group.^{1,2} Silicon reagents play important roles as not only protecting agents and masked nucleophiles but also Lewis acids. The Lewis acidity of alkyl- and alkoxy-silanes is much weaker than that of the corresponding group III metal compounds. However, when one of the four ligands on silicon is changed to a soft Lewis base such as triflate (OTf) or iodide (I), the acidity becomes strong enough for synthetic use as Lewis acid catalyst.³⁻⁸ Trimethylsilyl trifluoromethanesulfonate (TMSOTf) and iodotrimethylsilane (TMSI), bearing one Lewis-acidic co-ordination site, interact strongly to various heteroatoms, particularly oxygen, to activate the carbon-heteroatom bond. Although it is also known that five co-ordinated silicates and angle strained silanes have moderate Lewis acidity, these types of Lewis acids are only utilized to intramolecularly assist the addition of silylated nucleophiles as shown in the reactions of allylsilicates^{9,10} and enoxysilacyclobutane.¹¹

The synthetic reactions mediated by TMSOTf³⁻⁵ and TMSI⁶⁻⁸ have been extensively developed in the last two decades. The silicon-centred Lewis acids are frequently employed for the reactions of trimethylsilylated nucleophiles (TMSNu) with acetals ($R^1R^2C(OR^3)_2$) and carbonyl compounds ($R^1R^2C=O$), which usually proceed with a catalytic amount of TMSX (X = OTf, I) under mild conditions, forming C-heteroatom, C-C, and C-H bonds in high efficiency (Scheme 8.1). The plausible catalytic cycles A and B consist of three parts: polarization of the substrates by TMSX, nucleophilic attack of TMSNu to the resultant oxonium or carbenium ion intermediates, and dissociation into TMSX and the products. This chapter describes the synthetic utility of the TMSOTf- and TMSI-catalysed reactions.

Cycle A.



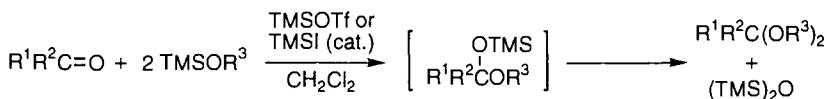
Cycle B.



Scheme 8.1

2. Carbon–oxygen and carbon–nitrogen bond formation

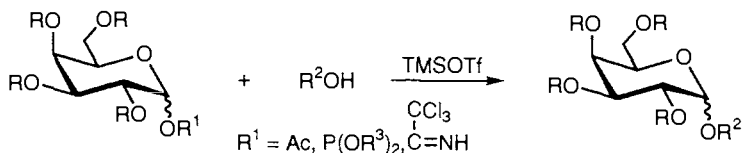
Aldehydes and ketones readily react with two equivalents of alkyl silyl ethers in the presence of TMSOTf or TMSI to afford dialkyl acetals in good yields (Scheme 8.2).^{12,13} The acid-catalysed reaction would initially form silyl alkyl acetals by nucleophilic addition of alkyl silyl ethers *via* the catalytic cycle B, and then, the acetals would turn dialkyl acetals by the successive nucleophilic substitution *via* the catalytic cycle A (Scheme 8.1). The driving force of the reaction is the great stability of hexamethyldisiloxane formed as a by-product. This method is particularly suited for the acetalization of unsaturated aldehydes and ketones without isomerization of the double bond.¹⁴



Scheme 8.2

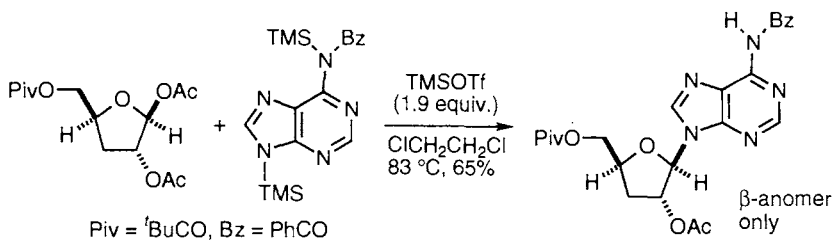
Transacetalization with alcohols or alkyl silyl ethers can also be induced by a catalytic or a stoichiometric amount of TMSOTf. The combination of TMSOTf with glycosyl acetates, trichloroacetimides, or phosphites is widely utilized as a tool of *O*-glycosylation for the synthesis of oligosaccharides (Scheme 8.3).^{15,16}

The prominent ability of TMSOTf is also exerted in nucleoside synthesis. The oxocarbenium ions generated from glycosyl acetates react with silylated heterocyclic bases to afford nucleosides (Scheme 8.4).^{17,18} In this case, an ex-



Scheme 8.3

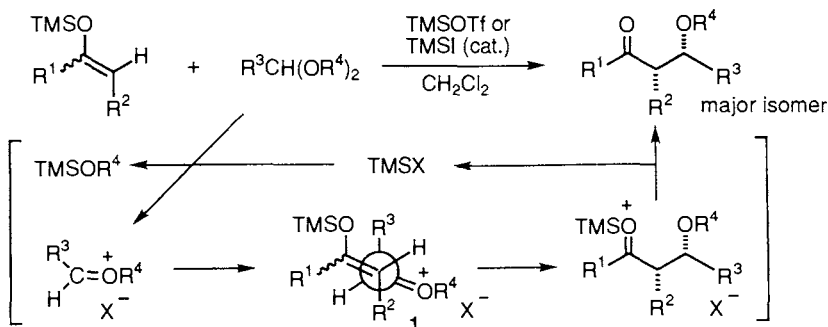
cess amount of TMSOTf is required for high efficiency. The major advantage of TMSOTf compared to SnCl₄ or other Lewis acids is to interact reversibly with the heterocycles without their deactivation.



Scheme 8.4

3. Carbon–carbon bond formation

The Mukaiyama cross-aldol reaction is one of the most powerful and selective methods for the construction of carbon–carbon bonds. This well-established transformation involves the reaction of silyl enolates with carbonyl compounds or their derivatives in the presence of an activator such as Lewis acid or fluoride ion.^{19,20} TMSOTf and TMSI can effectively promote this process with a catalytic quantity unlike TiCl₄, SnCl₄, and BF₃•Et₂O, which are required with a stoichiometric quantity. For instance, acetals and orthoesters react with silyl enolates in the presence of 1–10 mol% of TMSOTf or TMSI in CH₂Cl₂ at –78 °C to give β -alkoxy carbonyl compounds in high yield (Scheme 8.5 and



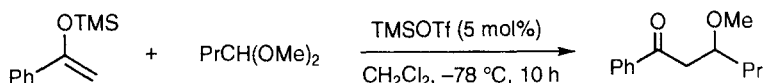
Scheme 8.5

Protocol 1).^{21–23} The aldol products do not undergo β -elimination of alcohols under the reaction conditions. In addition, this aldol reaction exhibits moderate to high erythroselectivity regardless of the geometry of silyl enolate.^{21,22} The stereochemical outcome can be rationalized by the acyclic transition state **1** in the reaction of silyl enolates with oxocarbenium ion intermediates. The aldol reaction of aldehydes and ketones with silyl enolates is also catalysed by TMSOTf,²⁴ although these electrophiles are less reactive than the corresponding acetals.

Protocol 1.

TMSOTf-catalysed reaction of 1-trimethylsiloxy-1-phenylethene with 1,1-dimethoxybutane^{21,22} (Scheme 8.6)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 8.6

Equipment

- A two-necked, round-bottomed flask (100 mL) fitted with a magnetic stirring bar, rubber septum, and three-way stopcock connected to a vacuum source and a nitrogen balloon
- Dry glass syringes (volume appropriate for quantity of solution to be transferred) with stainless-steel needles
- Dry ice–methanol bath

Materials

- 1,1-Dimethoxybutane^a (FW 118.2), 1.30 g, 11 mmol **flammable, irritant**
- 1-Trimethylsiloxy-1-phenylethene^b (FW 192.3), 1.92 g, 2.05 mL, 10 mmol **moisture sensitive, irritant**
- Trimethylsilyl trifluoromethanesulfonate^c (FW 222.3), 1.11 g, 0.97 mL, 0.50 mmol **flammable, corrosive**
- Freshly distilled (CaH₂) dichloromethane, 30 + 5 mL **toxic, irritant**

1. Dry the two-necked flask with an electric heat gun under vacuum, and introduce nitrogen gas from the balloon. Repeat this operation two more times. After cooling to room temperature, add 1,1-dimethoxybutane and 30 mL of dry CH₂Cl₂ *via* syringes.
2. Stir the mixture and cool to –78 °C (dry ice–methanol bath). Add 1-trimethylsiloxy-1-phenylethene to the solution *via* a syringe.
3. After 10 min, add a solution of trimethylsilyl trifluoromethanesulfonate in 5 mL of CH₂Cl₂ dropwise over 5 min *via* a syringe.
4. After stirring for 10 h at –78 °C, remove the septum and pour the reaction mixture into 50 mL of saturated aqueous NaHCO₃ solution.
5. Transfer the bilayer to a separating funnel and remove the organic layer. Extract the aqueous layer twice with 30 mL of CH₂Cl₂.

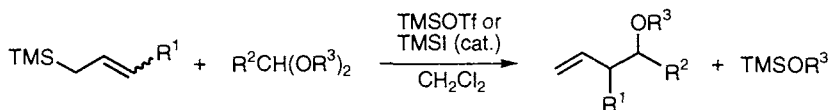
6. Pass the combined organic layer through a short K_2CO_3 column and concentrate *in vacuo*.
7. Purify the resulting oil by column chromatography on silica gel using hexane–ether (10:1). 3-Methoxy-1-phenylhexan-1-one (FW 206.3) is obtained in 75% yield (1.55 g).

^a Prepared by acid-catalysed reaction of butanal with methanol.²⁵

^b Prepared according to a published procedure.²⁶ Commercially available from Aldrich.

^c A colourless fluid liquid boiling at 45–47 °C at 17 mmHg. Commercially available from Aldrich. It can be easily prepared from trimethylsilyl chloride and trifluoromethane sulfonic acid.²⁷

The Hosomi–Sakurai reaction, the Lewis acid- or fluoride ion-mediated allylation using allylsilanes, can transfer functionalized allyl groups onto various carbon electrophiles with high regio-, stereo-, and chemoselectivity.²⁹ This synthetically valuable transformation as well as the Mukaiyama reaction is effectively induced by TMSOTf and TMSI. In the presence of a catalytic amount of TMSOTf or TMSI, acetals react with allylsilanes at the γ -position to give homoallyl ethers in good yields (Scheme 8.7).^{30–32} The use of chiral crotylsilane derivatives achieves high levels of acyclic stereocontrol.³³ Aldehydes and ketones can be directly led to homoallyl ethers by *in situ* acetalization with alkyl silyl ethers followed by allylation (Protocol 2).¹³ This consecutive reaction is applicable to asymmetric synthesis of homoallyl alcohols using homochiral alkyl silyl ethers.³⁴



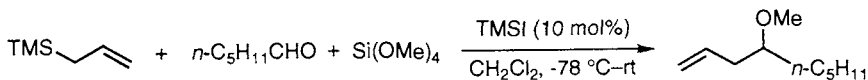
Scheme 8.7

Although TMSOTf and TMSI effect the allylation of aldehydes, their catalytic activities are considerably low. However, TMSB(OTf)₄, a supersilylating agent, is an excellent catalyst for the allylation.³⁵

Protocol 2.

TMSI-catalysed consecutive acetalization and allylation of hexanal to 4-methoxy-1-nonene¹³ (Scheme 8.8)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 8.8

Protocol 2. Continued**Equipment**

- A two-necked, round-bottomed flask (100 mL) fitted with a magnetic stirring bar, rubber septum, and three-way stopcock connected to a vacuum source and a nitrogen balloon
- Dry glass syringes (volume appropriate for quantity of solution to be transferred) with stainless-steel needles
- Dry ice–methanol bath and ice–water bath

Materials

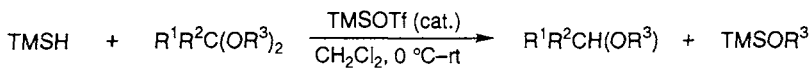
- | | |
|---|-----------------------------|
| • Freshly distilled (CaCl ₂) hexanal (FW 100.2), 1.00 g, 1.20 mL, 10 mmol | flammable, irritant |
| • Freshly distilled (CaH ₂) dichloromethane, 30 mL | toxic, irritant |
| • Tetramethoxysilane (FW 152.2), 1.67 g, 1.62 mL, 11 mmol | flammable, corrosive |
| • Allyltrimethylsilane (FW 114.3), 1.37 g, 1.91 mL, 12 mmol | flammable, irritant |
| • Iodotrimethylsilane ^a (FW 200.1), 0.20 g, 0.14 mL, 1.0 mmol | flammable, corrosive |
| • Pyridine, 0.5 mL | flammable, toxic |

1. Dry the two-necked flask with an electric heat gun under vacuum, and introduce nitrogen gas from the balloon. Repeat this operation two more times. After cooling to room temperature, add hexanal, dry CH₂Cl₂ tetramethoxysilane, and allyltrimethylsilane *via* syringes.
2. Stir the mixture and cool to –78°C (dry ice–methanol bath).
3. Add iodotrimethylsilane to the solution *via* a syringe and stir for 15 min at –78°C.
4. Then, warm the reaction vessel to 0°C with an ice–water bath.
5. After stirring for 3 h, add pyridine *via* a syringe.
6. After stirring for 5 min, remove the septum from the flask and pour the reaction mixture into 50 mL of saturated aqueous NaHCO₃ solution.
7. Transfer the bilayer to a separating funnel and remove the organic layer. Extract the aqueous layer twice with 30 mL of Et₂O.
8. Pass the combined organic layer through a short Na₂SO₄ column and concentrate *in vacuo*.
9. Purify the resulting oil by column chromatography on silica gel using hexane–ether (20:1). 4-Methoxy-1-nonen (FW 156.3) is obtained in 94% yield (1.47 g).

^aA colourless liquid boiling at 108°C at 760 mmHg. Commercially available. It can be easily prepared from hexamethyldisilane and iodine in quantitative yield as shown in Protocol 4.³⁶

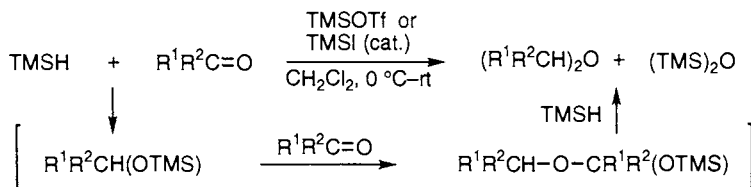
4. Carbon–hydrogen bond formation

A catalytic amount of TMSOTf successfully induces the reductive cleavage of acetals with trimethylsilane (TMSH) to afford the corresponding ethers in high yields (Scheme 8.9).³⁷ However, the TMSOTf- and TMSI-catalysed reactions of aldehydes or ketones with TMSH do not give silyl ethers but



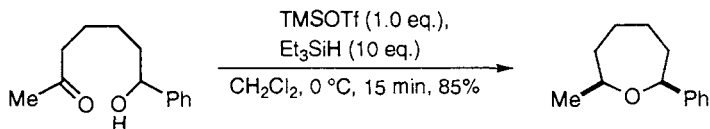
Scheme 8.9

symmetrical ethers (Scheme 8.10).³⁸ This reductive condensation can be applied for the cyclization of 1,4- and 1,5-diketones to cyclic ethers.³⁹ The formation of the symmetrical ethers is probably because the intermediary silyl ether attacks the activated C=O bond faster than TMSH. Indeed, under the similar conditions, the reaction of ketones or aldehydes with alkyl silyl ethers results in the selective formation of unsymmetrical ethers (Protocol 3).³⁸



Scheme 8.10

The combination of TMSOTf with hydrosilanes is also effective for the reductive cyclization of hydroxyketones (Scheme 8.11).⁴⁰

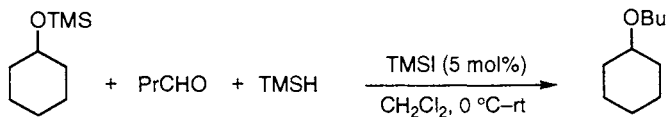


Scheme 8.11

Protocol 3.

TMSI-catalysed reductive condensation of butanal with cyclohexyloxytrimethylsilane using trimethylsilane³⁸ (Scheme 8.12)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 8.12

Protocol 3. Continued**Equipment**

- A three-necked, round-bottomed flask (100 mL) fitted with a magnetic stirring bar, a rubber septum, thermometer, and three-way stopcock connected to a vacuum source and a nitrogen balloon
- A bubbler containing liquid paraffin
- Tygon tubing attached to a stainless-steel needle
- Dry glass syringes (volume appropriate for quantity of solution to be transferred) with stainless-steel needles
- Ice–water bath

Materials

- Iodine (FW 253.8), 0.13 g, 0.50 mmol **highly toxic, corrosive**
- Hexamethyldisilane (FW 146.4), 79 mg, 0.11 mL, 0.54 mmol **flammable, irritant**
- Freshly distilled (CaH₂) dichloromethane, 14 + 10 mL **toxic, irritant**
- Freshly distilled (CaCl₂) butanal (FW 72.11), 0.72 g, 0.90 mL, 10 mmol **flammable, corrosive**
- Cyclohexyloxytrimethylsilane^a (FW 172.3), 1.72 g, 2.01 mL, 10 mmol **flammable, moisture sensitive**
- Trimethylsilane^b (FW 74.2) in a gas cylinder **flammable**

1. Dry the three-necked flask with an electric heat gun under vacuum, and introduce nitrogen gas from the balloon. Repeat this operation two more times. After cooling to room temperature, remove the septum, quickly add iodine, and attach the septum under a stream of nitrogen.
2. Add hexamethyldisilane and 14 mL of dry CH₂Cl₂ *via* syringes.
3. Stir the violet solution at room temperature for 10 min, and then, cool to 0°C with the ice–water bath.
4. Add a solution of butanal and cyclohexyloxytrimethylsilane in 10 mL of CH₂Cl₂ *via* a syringe.
5. After stirring for 10 min, replace the vacuum source with a bubbler and turn the stopcock to the bubbler from the nitrogen source.
6. Add trimethylsilane directly from a gas cylinder by means of Tygon tubing attached to a stainless-steel needle inserted through the septum. Slowly bubble the gas through the solution until the colour changes from violet to red–gold. During the bubbling, the internal temperature rises from 0°C to 15°C.
7. Stop the addition of trimethylsilane, remove the cold bath, and stir at room temperature for 2 h.
8. Remove the septum and pour the reaction mixture into 30 mL of 10% aqueous Na₂S₂O₃ solution.
9. Transfer the bilayer to a separating funnel and remove the aqueous layer. Similarly, wash the organic layer three more times with 30 mL of 10% aqueous Na₂S₂O₃ solution and four times with 30 mL of water.
10. Dry the organic layer over MgSO₄ and concentrate *in vacuo*.

11. Purify the crude product by distillation. Cyclohexyl butyl ether (FW 156.3, b.p. 68–70°C/0.08 mmHg) is obtained in 93% yield (1.45 g).

^aPrepared from cyclohexanol and chlorotrimethylsilane in the presence of triethylamine.⁴¹ Commercially available from Shin-Etsu Chemical Co. Ltd.

^bB.p. 6.7°C. Commercially available from Strem Chemicals Inc. or Chisso Co. Ltd. It can be prepared from chlorotrimethylsilane and lithium aluminum hydride.⁴²

References

1. Patai, S.; Rappoport, Z. *The Chemistry of Organic Silicon Compounds*; Wiley, Chichester, **1989**.
2. Colvin, E. W. *Silicon in Organic Synthesis*; Butterworths, London, **1981**.
3. Simchen, G. *Advances in Silicon Chemistry*. (ed G. L. Larson). JAI Press, Greenwich, **1991**, Vol. 1, p 189.
4. Noyori, R.; Murata, S.; Suzuki, M. *Tetrahedron* **1981**, 37, 3899.
5. Suzuki, M.; Noyori, R. *J. Syn. Org. Chem. Jpn* **1982**, 40, 534.
6. Olah, G. A.; Prakash, G. K. S.; Krishnamurti, R. *Advances in Silicon Chemistry* (ed G. L. Larson). JAI Press, Greenwich, **1991**, Vol. 1, p. 1.
7. Olah, G. A.; Narang, S. C. *Tetrahedron* **1982**, 38, 2225.
8. Hosomi, A. *J. Syn. Org. Chem., Jpn* **1982**, 40, 545.
9. Hosomi, A.; Kohra, S.; Tominaga, Y. *J. Chem. Soc., Chem. Commun*, **1987**, 1517.
10. Hosomi, A. *Rev. Heteroatom Chem.* **1992**, 7, 214.
11. Denmark, S. E.; Griedel, B. D.; Coe, D. M.; Schnute, M. E. *J. Am. Chem. Soc.* **1994**, 116, 7026.
12. Tsunoda, T.; Suzuki, M.; Noyori, R. *Tetrahedron Lett*, **1980**, 21, 1357.
13. Sakurai, H.; Sasaki, K.; Hayashi, J.; Hosomi, A. *J. Org. Chem.* **1984**, 49, 2808.
14. Seyferth, D.; Menzel, H.; Dow, A. W.; Flood, T. C. *J. Organomet. Chem.* **1972**, 44, 279.
15. Paulsen, H. *Angew. Chem., Int. Ed. Engl.* **1990**, 29, 823.
16. Lin, C.-C.; Shimazaki, M.; Heck, M.-P.; Aoki, S.; Wang, R.; Kimura, T.; Ritzèn, H.; Takayama, S.; Wu, S.-H.; Weitz-Schmidt, G.; Wong, C.-H. *J. Am. Chem. Soc.* **1996**, 118, 6826.
17. Vorbrüggen, H.; Kroliekiewicz, K.; Bennua, B.; *Chem. Ber.* **1982**, 114, 1234.
18. McDonald, F. E.; Gleason, M. M. *J. Am. Chem. Soc.* **1996**, 118, 6648.
19. Mukaiyama, T.; Narasaka, K.; Banno, K. *Chem. Lett.* **1973**, 1011.
20. Gennari, C. *Comprehensive Organic Synthesis* (ed. B. M. Trost, I. Fleming). Pergamon Press, Oxford, **1991**, Vol. 2, p 629.
21. Murata, S.; Suzuki, M.; Noyori, R. *J. Am. Chem. Soc.* **1980**, 102, 3248.
22. Murata, S.; Suzuki, M.; Noyori, R. *Tetrahedron* **1988**, 44, 4259.
23. Sakurai, H.; Sasaki, K.; Hosomi, A. *Bull. Chem. Soc. Jpn* **1983**, 56, 3195.
24. Mezger, F.; Simchen, G.; Fischer, P. *Synthesis* **1991**, 375.
25. Marsmann, H. C.; Horn, H.-G. *Z. Naturforsch., [b]* **1972**, 27, 1448.
26. House, H. O.; Czuba, L. J.; Gall, M.; Olmstead, H. D. *J. Org. Chem.* **1969**, 34, 2324.
27. DeGraw, J. I.; Goodman, L.; Baker, B. R. *J. Org. Chem.* **1961**, 26, 1156.
28. Hosomi, A.; Sakurai, H. *Tetrahedron Lett*, **1976**, 1295.
29. Hosomi, A. *Acc. Chem. Res.* **1988**, 21, 200.

30. Tsunoda, T.; Suzuki, M.; Noyori, R. *Tetrahedron Lett.* **1980**, 21, 71.
31. Sakurai, H.; Sasaki, K.; Hosomi, A. *Tetrahedron Lett.* **1981**, 22, 745.
32. Hosomi, A.; Sakata, Y.; Sakurai, H. *Tetrahedron Lett.* **1984**, 25, 2383.
33. Panek, J. S.; Yang, M. *J. Org. Chem.* **1991**, 56, 5755.
34. Tietze, L. F.; Schiemann, K.; Wegner, C. *J. Am. Chem. Soc.* **1995**, 117, 5851.
35. Davis, A. P.; Jaspars, M. *Angew. Chem. Int. Ed. Engl.* **1992**, 31, 470.
36. Kumada, M.; Shiina, A.; Yamaguchi, M. *Kogyo Kagaku Zasshi* **1954**, 57, 230.
37. Tsunoda, T.; Suzuki, M.; Noyori, R. *Tetrahedron Lett.* **1979**, 4679.
38. Sassaman, M. B.; Kotian, K. D.; Prakash, G. K. S.; Olah, G. A. *J. Org. Chem.* **1987**, 52, 4314.
39. Sassaman, M. B.; Prakash, G. K. S.; Olah, G. A. *Tetrahedron* **1988**, 44, 3771.
40. Nicolaou, K. C.; Hwang, C.-K.; Duggan, M. E.; Nugiel, D. A.; Abe, Y.; Reddy, K. B.; DeFrees, S. A.; Reddy, D. R.; Awartani, R. A.; Conley, S. R.; Rutjes, F. P. J. T.; Theodorakis, E. A. *J. Am. Chem. Soc.* **1995**, 117, 10227.
41. Corey, E. J.; Snider, B. B. *J. Am. Chem. Soc.* **1972**, 94, 2549.
42. Steward, O. W.; Pierce, O. R. *J. Am. Chem. Soc.* **1961**, 83, 1916.

Silver and gold reagents

MASAYA SAWAMURA

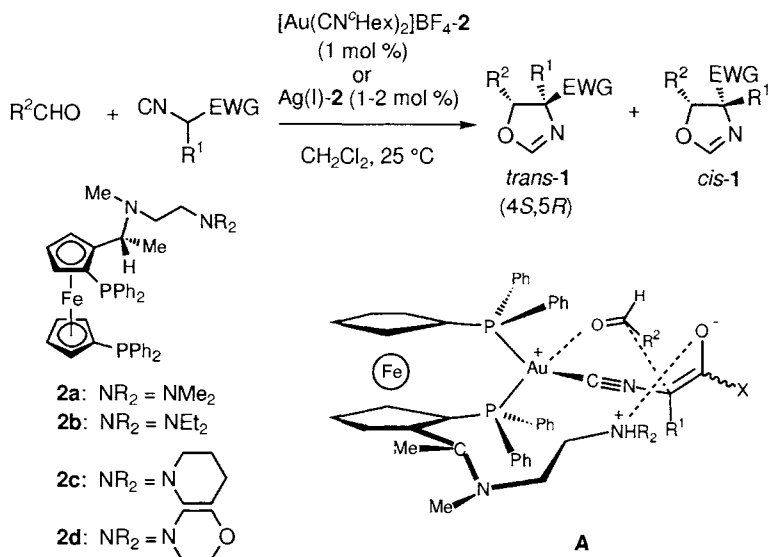
1. Introduction

Owing to their weak Lewis-acidity, reagents involving a silver(I) or gold(I) ion have attracted little attention as a Lewis-acid reagent to date. As described in this chapter, however, phosphine-co-ordinated silver(I) and gold(I) complexes have been successfully applied to the aldol-type condensation of isocyanides with aldehydes, where the silver(I) and gold(I) complexes activate the isocyanide through η^1 -co-ordination of the isocyano carbon (see Section 2). Although these complexes show rather low affinity to an oxygen functional group such as a carbonyl group, an example of carbonyl group activation can be seen in the silver(I)-catalysed allylation of aldehydes with allyltin reagents as a Lewis acid catalyst (see Section 3).

2. Enantioselective aldol reaction of activated isocyanides and aldehydes

Aldol-type reaction of activated isocyanides with aldehydes are effectively catalysed by a gold(I) or silver(I) complex of a chiral bis(phosphino)ferrocene bearing an ethylenediaminoalkyl pendant, producing optically active oxazolines.^{1,2} The activating group on the isocyanides includes an alkoxycarbonyl,^{3–20} carbamoyl,^{21–23} phosphonyl,^{24,25} or sulfonyl group²⁶ (Scheme 9.1, Tables 9.1 and 9.2). The reaction catalysed by silver(I) complex requires a slow addition of the isocyanide (over 1 h) except for the reaction of tosylmethyl isocyanide.

In most cases, the *trans*-oxazoline is predominantly formed with a high enantioselectivity, whereas the *cis*-oxazoline with low enantiomeric excess is as a minor product. The *trans*-oxazolines can readily be converted into *threo*- β -hydroxy- α -amino acids or their phosphonic acid analogues by acidic hydrolysis (Scheme 9.2). The reactions of α -substituted isocyanoacetate esters give, after acidic hydrolysis, α -alkylated α -amino acids.^{7,8} The oxazolines obtained from *N*-methoxy-*N*-methyl- α -isocyanocetamide (α -isocyano Weinreb amide)²³ can be transformed to *N,O*-protected β -hydroxy- α -amino aldehydes and ketones in high yield (Scheme 9.3).



Scheme 9.1

Table 9.1. Gold(I)-catalysed enantioselective aldol reaction of activated isocyanides with aldehydes^a

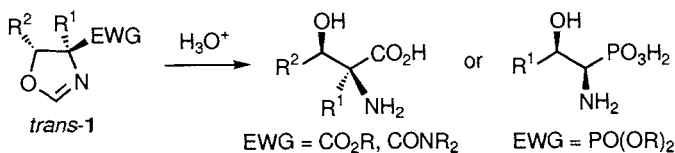
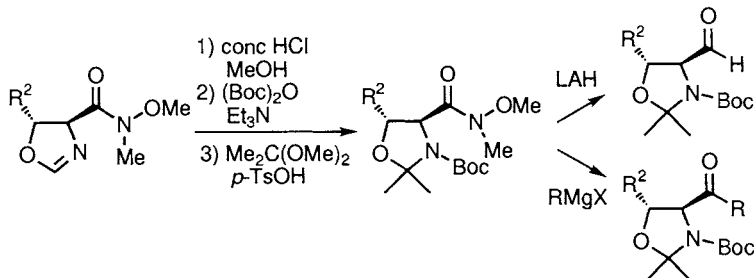
| EWG | R ¹ | R ² | Ligand | <i>trans</i> : <i>cis</i> | ee (%) (<i>trans</i>) |
|----------------------|----------------|--|-----------|---------------------------|-------------------------|
| CO ₂ Me | H | Ph | 2b | 89:11 | 93 |
| CO ₂ Me | H | Ph | 2c | 94:6 | 95 |
| CO ₂ Me | H | Me | 2a | 78:22 | 37 |
| CO ₂ Me | H | Me | 2d | 89:11 | 89 |
| CO ₂ Me | H | <i>i</i> -Pr | 2c | 99:1 | 94 |
| CO ₂ Me | H | <i>t</i> -Bu | 2d | >99:1 | 97 |
| CO ₂ Me | H | (<i>E</i>)-PrCH=CH | 2d | 87:13 | 92 |
| CO ₂ Me | Me | Ph | 2d | 93:7 | 94 |
| CO ₂ Me | <i>i</i> -Pr | Ph | 2c | 54:46 | 92 |
| CO ₂ Me | <i>i</i> -Pr | H | 2c | — | 81 |
| CONMe ₂ | H | 4-BnOC ₆ H ₄ CH ₂ | 2c | >95:5 | 95 |
| CON(Me)OMe | H | (<i>E</i>)-BnOCH ₂ CH=CH | 2d | 96:4 | 95 |
| PO(OPh) ₂ | H | Ph | 2c | >98:2 | 96 |

^a Yield is generally high.

The terminal amino group of the chiral ligand is essential for the catalyst activity and the stereoselectivities.³ Structure of the terminal amino group has a substantial effect on the stereoselectivity, six-membered ring amino groups such as piperidino or morpholino groups being generally efficient.^{3,4,16} The

Table 9.2. Silver(I)-catalysed enantioselective aldol reaction of activated isocyanides with aldehydes^a

| EWG | R ¹ | R ² | Ligand | <i>trans:cis</i> | ee (%) (<i>trans</i>) |
|---|----------------|----------------------|-----------|------------------|-------------------------|
| CO ₂ Me ^b | H | Ph | 2c | 96:4 | 80 |
| CO ₂ Me ^b | H | <i>i</i> -Pr | 2c | 99:1 | 90 |
| SO ₂ (<i>p</i> -Tol) ^c | H | (<i>E</i>)-MeCH=CH | 2c | 97:3 | 85 |

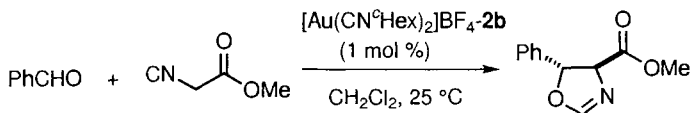
^a Yield is generally high.^b Reaction with 2 mol % of AgClO₄·2 under slow addition conditions.^c Reaction with 1 mol % of AgOTf·2.**Scheme 9.2****Scheme 9.3**

high efficiency of the gold catalysts has been explained by a transition state model as in structure **A**.^{15,27} The chiral ligand chelates to the gold atom with the two phosphorus atoms leaving the two nitrogen atoms unco-ordinated. The α -methylene protons of isocyanoacetate are activated through the coordination of the isocyano group to the gold atom, and the terminal amino group abstracts one of the activated α -protons, forming an ion pair between enolate anion and ammonium cation. The attractive ligand-substrate interaction permits a favourable arrangement of the enolate and aldehyde on the gold atom in the stereodifferentiating transition state.

Protocol 1.

Enantioselective aldol reaction of methyl isocyanoacetate with benzaldehyde catalysed by a chiral diamino bisphosphine-gold (I) complex³ (Scheme 9.4).

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 9.4

Equipment

- Magnetic stirrer
- Teflon-coated magnetic stirring bar (octagon 15 × 6.5 mm)
- Schlenk flask (30 mL)
- Syringes
- Thermostat
- Vacuum/inert gas source (argon)
- Water bath
- Kugelrohr distillation apparatus

Materials

- Bis(cyclohexyl isocyanide)gold(I) tetrafluoroborate,²⁸ 25.1 mg, 0.050 mmol air-sensitive
- (*R*)-*N*-Methyl-*N*-[2-(diethylamino)ethyl]-1-[(*S*)-1',2-bis(diphenylphosphino)ferrocenyl]ethylamine,⁵ 35.5 mg, 0.050 mmol
- Methyl isocyanoacetate,^a 0.495 g, 5.00 mmol stench, irritant, corrosive, lachrymator
- Benzaldehyde,^a 0.586 g, 5.50 mmol highly toxic, cancer suspect agent
- Dry, distilled dichloromethane, 5 mL toxic, irritant
- Dichloromethane for the transfer of a reaction mixture, 5 mL toxic, irritant

1. Place bis(cyclohexyl isocyanide)gold(I) tetrafluoroborate (25.1 mg, 0.050 mmol) and (*R*)-*N*-methyl-*N*-[2-(diethylamino)ethyl]-1-[(*S*)-1',2-bis(diphenylphosphino)ferrocenyl]ethylamine (35.5 mg, 0.050 mmol) in a 30 mL Schlenk flask with a magnetic stirring bar.
2. Purge the flask with argon.
3. Add dry distilled dichloromethane (5 mL) and stir the mixture at room temperature for 5 min.
4. Add methyl isocyanoacetate (0.495 g, 5.00 mmol) and immerse the flask into a water bath controlled at 25°C with a thermostat.
5. Add benzaldehyde (0.586 g, 5.50 mmol) and stir the mixture for 20 h. The completion of reaction can be checked by TLC (silica gel, hexane–ethyl acetate, 2:1).

6. Remove the solvent under reduced pressure.
7. Distil the residual liquid with a Khugelrohr (ca. 100°C, 1 mmHg) to obtain 1.05 g (98%) of the oxazoline or **3a** as a mixture of *trans*-(4*S*,5*R*) (93% ee) and *cis*-(4*R*,5*R*) (49% ee) isomers in a ratio of 89:11 (¹H NMR).
8. Determine the enantiomeric excesses of both isomers by ¹H NMR analysis with a chiral shift reagent Eu(hfc)₃^b or by GLC analysis with a chiral stationary phase capillary column (Sumichiral OA-520^c).
9. The *trans*- and *cis*-isomers can readily be separated by flash column chromatography on silica gel (hexane–ethyl acetate, 5:1).^d

^a Purify by distillation before use.

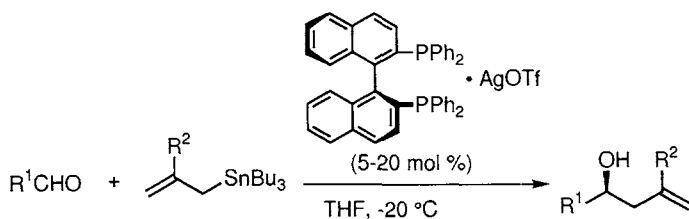
^b Commercially available.

^c Sumitomo Chemical Co. (Japan).

^d Pre-dry the solvent over CaCl₂ and use the supernatant.

3. Enantioselective aldehyde allylation with allyltin reagents catalysed by a silver(I)–phosphine complex

The cationic silver(I)-BINAP complex is an efficient catalyst for the enantioselective allylation of aldehydes with allyltin reagents (Scheme 9.5).²⁹ The catalyst is prepared *in situ* from silver(I) triflate and the BINAP ligand in THF, and typically the reaction is carried out in the same solvent at –20°C. High enantioselectivity is observed with various aldehydes including aromatic, α,β-unsaturated, and saturated aliphatic aldehydes, whereas the scope for the saturated aliphatic aldehydes are somewhat limited (Table 9.3).



Scheme 9.5

Mechanistic studies suggest that the silver(I)-BINAP complex acts as a Lewis acid catalyst rather than an allylsilver reagent.

Table 9.3. Enantioselective allylation of aldehydes with allyltin reagents catalysed by a silver(I) complex of (*S*)-BINAP^a

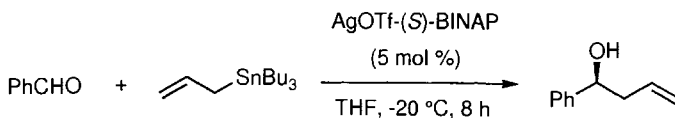
| R ¹ | R ² | S/C ^b | Yield (%) | ee (%) |
|-----------------------------------|----------------|------------------|-----------|------------------------------|
| Ph | H | 20 | 88 | 96 (<i>S</i>) |
| Ph | Me | 20 | 75 | 92 (<i>R</i>) ^c |
| 2-MeC ₆ H ₄ | H | 20 | 85 | 97 |
| 4-MeC ₆ H ₄ | H | 20 | 59 | 97 |
| (<i>E</i>)-PhCH=CH | H | 6.7 | 83 | 88 (<i>S</i>) |
| PhCH ₂ CH ₂ | H | 5 ^d | 47 | 88 |

^a Allyltin reagent:aldehyde, 1–4:1.^b Aldehyde:catalyst.^c Reaction with (*R*)-BINAP.^d The reaction was started with 0.1 equiv of the catalyst, and 0.1 equiv of the catalyst was added after 4 h.

Protocol 2.

Asymmetric allylation of benzaldehyde with allyltributyltin catalysed by (*S*)-BINAP·AgOTf Complex. Preparation of (*S*)-1-Phenyl-3-butene-1-ol²⁹ (Scheme 9.6)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 9.6**

Equipment

- Magnetic stirrer
- Syringes
- Vacuum/inert gas source (argon)
- Schlenk flask (20 mL)
- Syringe pump
- Cooling bath with dry ice–*o*-xylene

Materials

- Silver trifluoromethanesulfonate,^a 26.4 mg, 0.103 mmol irritant, light-sensitive
- (*S*)-(–)-2,2'-Bis(diphenylphosphino)-1,1'-binaphthyl,^a 66.5 mg, 0.107 mmol
- Benzaldehyde,^b 208.0 mg, 1.96 mmol highly toxic, cancer suspect agent
- Allyltributyltin,^b 663.1 mg, 2.00 mmol irritant
- Dry tetrahydrofuran (THF),^c 6 mL flammable, irritant
- 1 M HCl solution in water, 10 mL toxic, corrosive
- Potassium fluoride, 1 g toxic, corrosive
- Magnesium sulfate hygroscopic

1. Ensure that all glass equipment has been dried in an oven before use.
2. Place silver trifluoromethanesulfonate (26.4 mg, 0.103 mmol) and (S)-(-)-2,2'-bis(diphenylphosphino)-1,1'-binaphthyl (66.5 mg, 0.107 mmol) in a 20 mL Schlenk flask with a magnetic stirring bar under argon atmosphere and exclusion of direct light. Add dry tetrahydrofuran (3 mL) and stir the mixture at 20°C for 10 min.
3. After cooling the resulting solution to -20°C, add benzaldehyde (208.0 mg, 1.96 mmol) dropwise. Then add a THF solution (3 mL) of allyltributyltin (663.1 mg, 2.00 mmol) over a period of 4 h with a syringe pump at -20°C.
4. After stirring for 4 h at this temperature, treat the reaction mixture with a mixture of 1 M HCl (10 mL) and solid KF (1 g) at ambient temperature for 30 min.
5. Filter off the resulting precipitate, dry the filtrate over MgSO₄ and concentrate *in vacuo*.
6. Purify the crude product by flash column chromatography on silica gel [hexane-ethyl acetate (10:1)] to give the (S)-enriched product (258 mg, 88% yield, 96% ee) as a colourless oil which displays the appropriate ¹H NMR (in CDCl₃) and IR (neat).
7. Determine the enantiomeric excess by HPLC analysis (Daicel Chiralcel OD-H, hexane-*i*-PrOH, 20:1; flow rate, 0.5 mL/min), *t_R* = 18.2 min (*R*-isomer), *T_R* = 19.9 min (*S*-isomer).

^a Commercially available products can be used without purification. Do not use aged ones.

^b Purify benzaldehyde and allyltributyltin by distillation before use.

^c Commercially available anhydrous THF (Aldrich) can be used as received.

References

1. Sawamura, M.; Ito, Y. *Chem. Rev.* **1992**, 92, 857.
2. Sawamura, M.; Ito, Y. In *Catalytic Asymmetric Synthesis* (ed. I. Ojima). VCH Publishers, New York, **1993**, p 367.
3. Ito, Y.; Sawamura, M.; Hayashi, T. *J. Am. Chem. Soc.* **1986**, 108, 6405.
4. Ito, Y.; Sawamura, M.; Hayashi, T. *Tetrahedron Lett.* **1987**, 28, 6215.
5. Hayashi, T.; Sawamura, M.; Ito, Y. *Tetrahedron*, **1992**, 48, 1999.
6. Ito, Y.; Sawamura, M.; Hayashi, T. *Tetrahedron Lett.* **1988**, 29, 239.
7. Ito, Y.; Sawamura, M.; Shirakawa, E.; Hayashizaki, K.; Hayashi, T. *Tetrahedron Lett.* **1988**, 29, 235.
8. Ito, Y.; Sawamura, M.; Shirakawa, E.; Hayashizaki, K.; Hayashi, T. *Tetrahedron* **1988**, 44, 5253.
9. Soloshonok, V. A.; Hayashi, T. *Tetrahedron Lett.* **1994**, 35, 2713.
10. Soloshonok, V. A.; Kacharov, A. D.; Hayashi, T. *Tetrahedron* **1996**, 52, 245.
11. Soloshonok, V. A.; Hayashi, T.; Ishikawa, K.; Nagashima, N. *Tetrahedron Lett.* **1994**, 35, 1055.

12. Soloshonok, V. A.; Kacharov, A. D.; Avilov, D. V.; Hayashi, T. *Tetrahedron Lett.* **1996**, 37, 7845.
13. Hayashi, T.; Uozumi, Y.; Yamazaki, A.; Sawamura, M.; Hamashima, H.; Ito, Y. *Tetrahedron Lett.* **1991**, 32, 2799.
14. Pastor, S. D.; Togni, A. *J. Am. Chem. Soc.* **1989**, 111, 2333.
15. Togni, A.; Pastor, S. D. *J. Org. Chem.* **1990**, 55, 1649.
16. Pastor, S. D.; Togni, A. *Helv. Chim. Acta* **1991**, 74, 905.
17. Togni, A.; Pastor, S. D. *Helv. Chim. Acta* **1989**, 72, 1038.
18. Togni, A.; Pastor, S. D.; Rihs, G. *Helv. Chim. Acta* **1989**, 72, 1471.
19. Togni, A.; Häusel, R. *Synlett* **1990**, 633.
20. Gorla, F.; Togni, A.; Venanzi, L. M.; Albinati, A.; Lianza, F. *Organometallics* **1994**, 13, 1607.
21. Ito, Y.; Sawamura, M.; Kobayashi, M.; Hayashi, T. *Tetrahedron Lett.* **1988**, 29, 6321.
22. Soloshonok, V. A.; Hayashi, T. *Tetrahedron Asymmetry* 1994, 5, 1091.
23. Sawamura, M.; Nakayama, Y.; Kato, T.; Ito, Y. *J. Org. Chem.* **1995**, 60, 1727.
24. Sawamura, M.; Ito, Y.; Hayashi, T. *Tetrahedron Lett.* **1989**, 30, 2247.
25. Togni, A.; Pastor, S. D. *Tetrahedron Lett.* **1989**, 30, 1071.
26. Sawamura, M.; Hamashima, H.; Ito, Y. *J. Org. Chem.* **1990**, 55, 5935.
27. Sawamura, M.; Ito, Y.; Hayashi, T. *Tetrahedron Lett.* **1990**, 31, 2723.
28. Bonati, F.; Minghetti, G. *Gazz. Chim. Ital.* **1973**, 103, 373.
29. Yanagisawa, A.; Nakashima, H.; Ishiba, A.; Yamamoto, H. *J. Am. Chem. Soc.* **1996**, 118, 4723.

Zr- and Hf-centred Lewis acid in organic synthesis

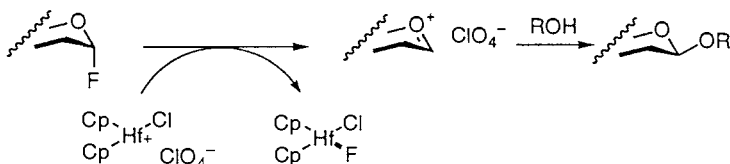
KEISUKE SUZUKI

1. Introduction

The nature of a given Lewis acid, ML_n , can be modified by changing the metal (M), the ligand (L, chiral or non-chiral) as well as the charge (neutral or cationic), thereby endowing it with useful properties in organic synthesis. Early transition metals are attracting current attention to exploit their hard Lewis acidic characters. Although titanium has been studied most, recent attention is also centred at its group-4 congeners, i.e. zirconium and hafnium. Although they have been extensively used in the polymerization field,¹⁻³ the use in organic synthesis is still rather limited either as a tailor-made catalyst or as a transient species in the course of the reaction.⁴ This chapter describes some unique reactivities of cationic zirconocene or hafnocene complexes, thereby dealing with four modes of functional group activation: (1) C-F bond in glycosyl fluoride, (2) ethers, (3) carbonyl groups, (4) C-C multiple bonds.

2. Activation of C-F bond

Combination of Cp_2HfCl_2 and $AgClO_4$ acts as a powerful activator of glycosyl fluoride, providing a glycosylation method used for the synthesis of complex oligosaccharides or other glycoconjugates (Scheme 10.1).⁵⁻⁸ The high fluorophilicity of hafnocene perchlorate complex, cationic, or loosely bound covalent species, is invoked to the origin of high activation of C-F bond to generate an oxonium species. Double ligand exchange by Cp_2HfCl_2 and $AgClO_4$ in 1:2



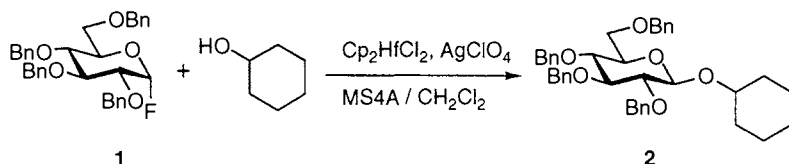
Scheme 10.1

ratio leads to even higher reactivity. Triflate is used as an alternative counter anion considering the potential hazard associated with perchlorate.

Protocol 1.

Metalloocene-promoted *O*-glycosylation (Scheme 10.2)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 10.2

Equipment

- Two-necked, round-bottomed flask (30 mL) containing a magnetic stirring bar, with a side arm bearing a three-way stopcock, and a rubber septum
- Vacuum/inert gas source (gas source may be an argon balloon)
- Dry ice–methanol cooling bath
- All-glass syringe

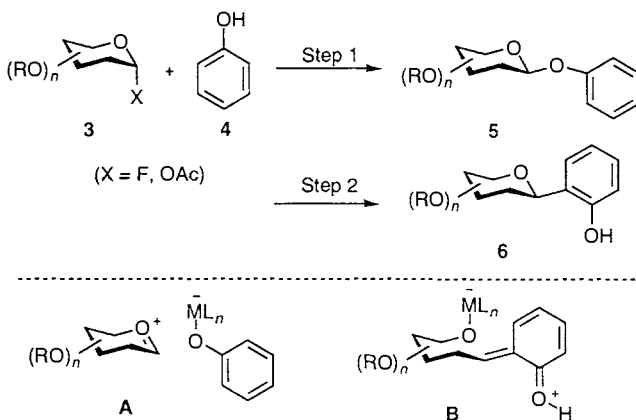
Materials

- AgClO_4 (FW 207), 47.0 mg, 0.227 mmol
 - Cp_2HfCl_2 (FW 379), 43.0 mg, 0.113 mmol
 - dry, distilled CH_2Cl_2
 - Powdered molecular sieves 4A, c. 100 mg
 - Cyclohexanol (FW 100), 22.6 mg, 0.226 mmol
 - Glycosyl fluoride **1** (FW 543), 61.2 mg, 0.113 mmol
- oxidizer, corrosive, explosive
irritant, moisture sensitive**
- irritant, hygroscopic**

1. Assemble the two-necked flask with a stopcock, a stirring bar, and a rubber septum. Flame dry the flask under vacuum, and, after cooling to ambient temperature, purge with argon through the three-way stopcock. Add Cp_2HfCl_2 and AgClO_4 to the flask.
2. Add dry CH_2Cl_2 (0.5 mL) via a syringe, and stir the mixture for 10 min to make a slurry.
3. Add a solution of cyclohexanol in CH_2Cl_2 (0.5 mL). After cooling to -50°C , add glycosyl fluoride **1** in CH_2Cl_2 (1 mL) via a syringe, and stir the mixture for 1 h.
4. Quench the reaction with saturated aqueous NaHCO_3 solution, and filter the mixture through a Celite pad. Extract the products with EtOAc, and wash the combined organic extracts successively with saturated aqueous NaHCO_3 solution and brine. After drying over Na_2SO_4 , and evaporate the solvents *in vacuo*.
5. Purify the resulting oil on preparative TLC (hexane–EtOAc, 4:1) to yield *O*-glycoside **2** (65.2 mg, 93%). The product shows characteristic spectroscopic data.

3. Activation of ethers

Cationic hafnocene species are also capable of activating the ether linkages as seen in the unique results in *C*-glycoside synthesis.⁹ In the *C*-glycosylation of phenol via *in situ*-formed *O*-glycoside promoted by a Lewis acid (**3** + **4** → [**5**] → **6**: *O* → *C* glycoside rearrangement), the Cp₂HfCl₂–AgClO₄ combination is particularly effective in terms of the reactivity in converting **5** into **6** as well as the stereoselectivity. Both of these features have their origin in the effective activation of an ether oxygen by a cationic hafnocene species as illustrated in **A** and **B**^{10–13} (Scheme 10.3).



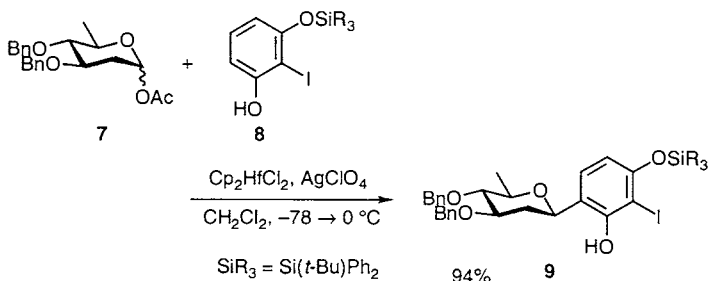
Scheme 10.3

Wipf *et al.*^{14–16} reported the activation of epoxide by a cationic zirconocene species to trigger 1,2-rearrangement.

Protocol 2.

O → *C*-glycoside rearrangement of olivosyl acetate **7** and phenol **8** (Scheme 10.4)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 10.4

Protocol 2. Continued**Equipment**

- Two-necked, round-bottomed flask (100 mL) fitted with a magnetic stirring bar, a three-way stopcock, and a rubber septum
- Dry ice–methanol cooling bath
- Vacuum/inert gas source (gas source may be an argon balloon)
- All-glass syringe

Materials

- AgClO₄ (FW 207.3), 893 mg, 4.31 mmol
- Cp₂HfCl₂ (FW 380), 820 mg, 2.16 mmol
- Dry, distilled CH₂Cl₂
- Powdered molecular sieves 4A, ca. 1.2 g
- Phenol **8** (FW 474), 936 mg, 1.97 mmol
- D-oliviosyl acetate **7** (FW 370), 666 mg, 1.80 mmol

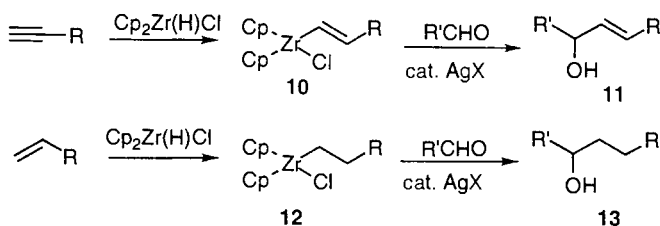
**oxidizer, corrosive, explosive
irritant, moisture sensitive**

1. Assemble the 100 mL round-bottomed flask with a stopcock, a stirring bar, and a rubber septum. Flame dry the vessel containing powdered molecular sieves under vacuum. After cooling, purge with argon. Place Cp₂HfCl₂ and AgClO₄ in the flask.
2. Introduce dry CH₂Cl₂ (5 mL) via a syringe through the rubber septum, and form a slurry by stirring for 15 min.
3. After cooling the mixture to −78°C, add a solution of phenol **8** in CH₂Cl₂ (15 mL) and D-oliviosyl acetate **7** in CH₂Cl₂ (10 mL) via a syringe to this suspension.
4. Let the mixture to warm to 0°C during 40 min, and stir it for 15 min.
5. Quench the reaction with pH 7 phosphate buffer, and acidify it with 2 M HCl. Filter the mixture through a Celite pad, and extract the products with EtOAc. Wash the combined organic extracts with brine. After drying over Na₂SO₄ and filtration, concentrate the solution *in vacuo*.
6. Purify the resulting oil on silica-gel flash column (Hexane–acetone, 85:15) to yield C-glycoside **9** (1.32 g, 93.5%). The product shows characteristic spectroscopic data.

4. Carbonyl activation

Carbonyl groups are also activated by cationic zirconocene species. Alkenyl- or alkyl-zirconocene chlorides, **10** or **12** (Scheme 10.5), readily accessible via hydrozirconation of alkyne or alkene,^{17,18} are unreactive nucleophiles, and their Grignard-type carbonyl addition is slow. However, a catalytic amount of AgClO₄ or AgAsF₆ greatly accelerates the reaction, which is ascribed to the carbonyl activation by the cationic species.¹⁹ Recently, Hf(OTf)₄ was prepared and used as a catalyst for Friedel–Crafts reaction, and for Fries

rearrangement, which are also ascribed to the carbonyl activation by the hafnium reagent^{20,21}

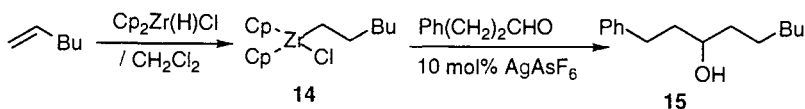


Scheme 10.5

Protocol 3.

AgAsF_6 catalysed addition of alkenylzirconocene chloride to aldehyde (Scheme 10.6)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 10.6

Equipment

- Two-necked, round-bottomed flask (30 mL) fitted with a magnetic stirring bar, a three-way stopcock, and a rubber septum
- Nitrogen gas line
- Vacuum/inert gas source (gas source may be an argon balloon)
- Dry ice–methanol cooling bath
- All-glass syringe

Materials

- $\text{Cp}_2\text{Zr(H)Cl}$ (FW 258), 270 mg, 1.05 mmol
 - AgAsF_6 (FW 296), 20.1 mg, 0.0677 mmol
 - Dry, distilled CH_2Cl_2
 - 1-hexyne (FW 84.2), 93.1 mg, 1.11 mmol
 - 3-phenylpropanal (FW 134), 83.1 mg, 0.619 mmol
- moisture sensitive, light sensitive**
highly toxic, cancer suspect agent

flammable liquid irritant

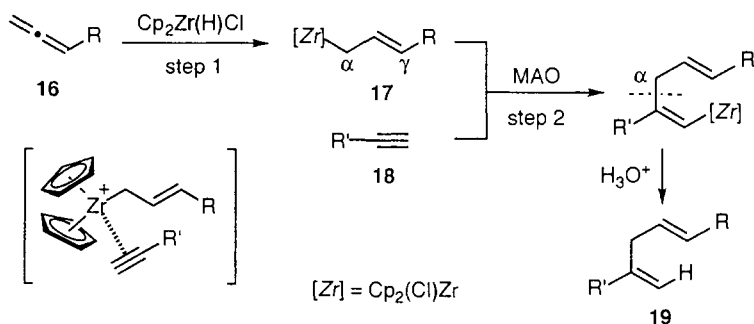
- Assemble the 30 mL round-bottomed flask with a stopcock, a stirring bar, and a rubber septum, and flame dry the flask under vacuum. After cooling to room temperature, fill the apparatus with nitrogen.
- Place Schwartz reagent in the flask, and then rapidly place the rubber septum on the neck of the flask.
- Add a solution of the 1-hexyne in CH_2Cl_2 (3 mL) via a syringe to the flask at -78°C , and then allow the mixture to warm to 25°C , and stir the mixture for 10 min.

Protocol 3. Continued

4. Add a solution of 3-phenylpropanal in CH_2Cl_2 (3 mL) to the mixture.
5. After 5 min, add AgAsF_6 to the mixture, where a deep brown suspension will form.
6. After 10 min, pour the mixture into saturated aqueous NaHCO_3 solution, extract the products with EtOAc. Wash the combined extracts with brine and dry over Na_2SO_4 .
7. Purify the resulting oil on preparative TLC (Hexane–EtOAc, 4:1) to yield **15** as colourless oil (129 mg, 94.5%). The product shows characteristic spectroscopic data.

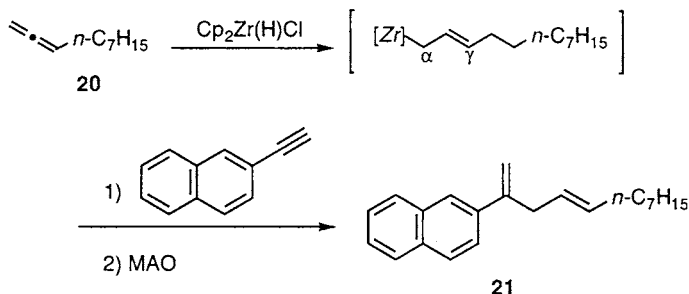
5. Reductive coupling of allenes and alkynes

It is well documented that cationic zirconocene species plays a key role in Kaminsky polymerization, a recent descendant of the Ziegler–Natta process.^{1–3} This process served as a hint to a regio- and stereocontrolled reductive coupling of allenes and alkynes (Scheme 10.7); allylzirconium species **17**, generated by the hydrozirconation of allene **16**, undergoes carbometallation of alkyne **18** promoted by methylaluminoxane (MAO). The overall reaction gives the ‘ α -internal’ coupling product **19** in high selectivity.²²

**Scheme 10.7**

Protocol 4.**MAO-catalysed allylzirconation of 1-alkynes (Scheme 10.8)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 10.8****Equipment**

- Two-necked, round-bottomed flask (30 mL) fitted with a magnetic stirring bar, a three-way stopcock, and a rubber septum
- Nitrogen gas line
- Vacuum/inert gas source (gas source may be an argon balloon)
- Dry ice-methanol cooling bath
- All-glass syringe

Materials

- $\text{Cp}_2\text{Zr}(\text{H})\text{Cl}$ (FW 258), 190 mg, 0.736 mmol
- Methylaluminoxane (MAO) (0.93 M solution in toluene), 0.12 mL
- Dry, distilled CH_2Cl_2
- Allene **20** (FW 138), 86.3 mg, 0.625 mmol
- 2-Ethynynaphthalene (FW 152), 55.4 mg, 0.364 mmol

moisture sensitive, light sensitive**pyrophoric, corrosive**

1. Assemble the two-necked flask with a stopcock, stirring bar, and a rubber septum, and flame dry the flask under vacuum. After cooling to room temperature, purge the flask with nitrogen.
2. Transfer Schwartz reagent from Schlenk flask into the flask which is connected to the nitrogen line, and then rapidly equip the neck of the flask with a rubber septum.
3. Add CH_2Cl_2 (0.7 mL) to the flask to form a slurry at -78°C . Add a solution of allene **20** in CH_2Cl_2 (2.5 mL), warm the mixture to 25°C over 20 min and stir for a further 0.5 h.
4. Chill the resulting red solution to -78°C , and add ethynynaphthalene in CH_2Cl_2 (2.5 mL) via a syringe followed by MAO solution. Warm the mixture to -20°C over 15 min, and stir for 20 min at -20°C . Stop the reaction by carefully adding saturated aqueous K_2CO_3 solution. Stir the mixture for 5 min, and add anhydrous Na_2SO_4 .

Protocol 4. Continued

5. After filtration through a Celite pad and evaporation, purify the resulting oil on preparative TLC (hexane) to give diene **21** (105 mg, 99%). The product shows characteristic spectroscopic data.
-

References

1. Brintzinger, H. H.; Fischer, D.; Mülhaupt, R.; Rieger, B.; Waymouth, R. M. *Angew. Chem. Int. Ed. Engl.* **1995**, *34*, 1143.
2. Jordan, R. F. *Adv. Organomet. Chem.* **1991**, *32*, 325.
3. Sinn, H.; Kaminsky, W. *Adv. Organomet. Chem.* **1980**, *18*, 99.
4. Suzuki, K. *Pure Appl. Chem.* **1994**, *66*, 1557.
5. Matsumoto, T.; Maeta, H.; Suzuki, K.; Tsuchihashi, G. *Tetrahedron Lett.* **1988**, *29*, 3567.
6. Suzuki, K.; Maeta, H.; Matsumoto, T.; Tsuchihashi, G. *Tetrahedron Lett.* **1988**, *29*, 3571.
7. Suzuki, K.; Maeta, H.; Matsumoto, T. *Tetrahedron Lett.* **1989**, *30*, 4853.
8. Matsumoto, T.; Maeta, H.; Suzuki, K.; Tsuchihashi, G. *Tetrahedron Lett.* **1988**, *29*, 3575.
9. Suzuki, K.; Matsumoto, T. Total synthesis of aryl C-glycoside antibiotics. In *Recent progress in the chemical synthesis of antibiotics and related microbial products*, Vol. 2 (ed. G. Lukacs) Springer Verlag, Berlin, pp 352–403.
10. Matsumoto, T.; Katsuki, M.; Suzuki, K. *Tetrahedron Lett.* **1988**, *29*, 6935.
11. Matsumoto, T.; Hosoya, T.; Suzuki, K. *Synlett* **1991**, 709.
12. Matsumoto, T.; Hosoya, T.; Suzuki, K. *Tetrahedron Lett.* **1988**, *31*, 4629.
13. Matsumoto, T.; Katsuki, M.; Jona, H.; Suzuki, K. *J. Am. Chem. Soc.* **1991**, *113*, 6982.
14. Wipf, P.; Xu, W. *J. Org. Chem.* **1993**, *58*, 825.
15. Wipf, P.; Xu, W. *J. Org. Chem.* **1993**, *58*, 5880.
16. Wipf, P.; Xu, W. *Tetrahedron* **1995**, *51*, 4551.
17. Schwartz, J.; Labinger, J. A. *Angew. Chem., Int. Ed. Engl.* **1976**, *15*, 333.
18. Buchwald, S. L.; LaMaire, S. J.; Nielsen, R. B.; Watson, B. T.; King, S. M. *Org. Synth.* **1992**, *71*, 77.
19. Suzuki, K.; Hasegawa, T.; Imai, T.; Maeta, H.; Ohba, S. *Tetrahedron*, **1995**, *55*, 4483.
20. Hachiya, I.; Moriwaki, M.; Kobayashi, S. *Bull. Chem. Soc. Jpn* **1995**, *68*, 2053.
21. Kobayashi, S.; Moriwaki, M.; Hachiya, I. *Bull. Chem. Soc. Jpn* **1997**, *70*, 267.
22. Yamanoi, S.; Imai, T.; Matsumoto, T.; Suzuki, K. *Tetrahedron Lett.* **1997**, *38*, 3031.

Scandium(III) and yttrium(III)

SHIN-ICHI FUKUZAWA

1. Introduction

According to the inorganic chemistry text book written by Cotton and Wilkinson,¹ scandium is not truly a rare earth element but yttrium and lanthanides are rare earth elements. The stable oxidation state of scandium is trivalent and the ionic radius of scandium (III) is significantly smaller (0.89 \AA) than those for any of the rare earth elements ($1.0\text{--}1.17 \text{ \AA}$). Chemical behaviour is intermediate between aluminium and that of lanthanides. Yttrium has a trivalent oxidation state similar to scandium and lanthanide elements and the ionic radius of yttrium (III) (1.04 \AA) is close to those of erbium (1.03 \AA) and holmium (1.04 \AA). Yttrium resembles lanthanide elements in its chemical properties.

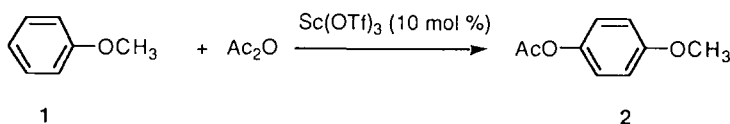
2. Scandium(III) triflate

Among scandium(III) compounds, scandium(III) trifluoromethanesulfonate (triflate) $[\text{Sc}(\text{OTf})_3]$ is the most attractive reagent and has been intensively studied in organic synthesis because it has a stronger Lewis acidity than that of other scandium (III) compounds.² Compared to even yttrium(III) and lanthanide(III) triflates, $\text{Sc}(\text{OTf})_3$ has a stronger Lewis acidity and catalyses certain reactions which are not mediated by yttrium(III) and lanthanide(III) triflates. The stronger Lewis acidity of $\text{Sc}(\text{OTf})_3$ is probably due to its smaller ionic radius than those of lanthanides. $\text{Sc}(\text{OTf})_3$ is commercially available or readily prepared by the reaction of scandium oxide and trifluoromethanesulfonic acid in water.³ Characteristic features of scandium triflate are as follows. It can be used as a catalyst in the presence of water and catalyses certain reactions and maintains catalytic activity. It can be recovered from the aqueous layer after reaction and reused in further reactions. As it is stable in water and air, it is easy to handle and the catalysed reaction can be carried out by a simple procedure.

2.1 Friedel–Crafts reaction

$\text{Sc}(\text{OTf})_3$ catalyses Friedel–Crafts acylation reaction of arenes effectively.⁴ It mediates the reaction more efficiently than $\text{Yb}(\text{OTf})_3$ and $\text{Y}(\text{OTf})_3$. Acylation

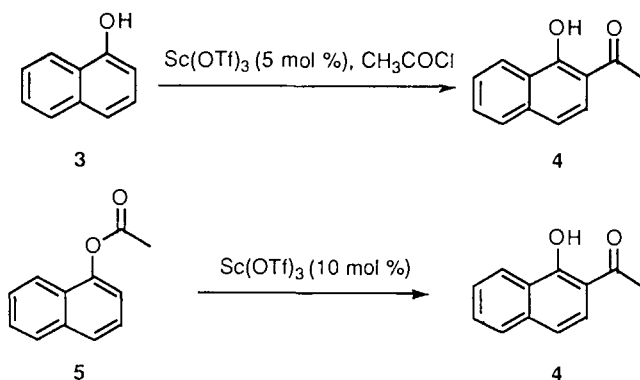
of benzene and toluene does not proceed, but reaction with more electron-rich arenes gives acylated products. Reaction with anisole, thioanisole, *o*- and *m*-dimethoxybenzene gives the corresponding single acylated product in excellent yields, respectively (Scheme 11.1).



Scheme 11.1

The advantages of the process are as follows. First, a catalytic amount (1 mol %) of $\text{Sc}(\text{OTf})_3$ is enough to promote the reaction in contrast to the conventional acylation reaction which requires a stoichiometric amount of Lewis acid. Second, evolution of hydrogen halide during quenching of Lewis acid, such as AlCl_3 , is avoided. Third, $\text{Sc}(\text{OTf})_3$ can be recovered from the aqueous layer and re-used as a catalyst, whereas disposal of a considerable amount of aluminium hydroxide which results from the work-up process is a severe environmental problem in the conventional Friedel–Crafts process.

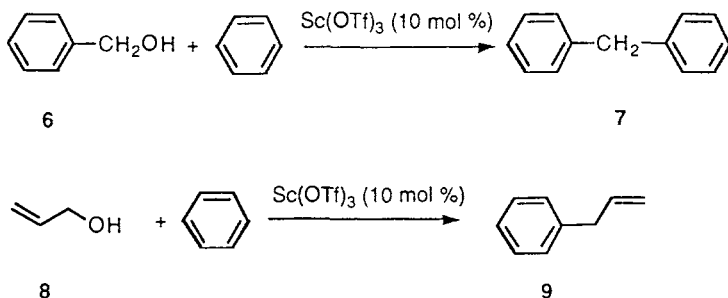
$\text{Sc}(\text{OTf})_3$ has been shown to catalyse *ortho*-selective acylation of phenol and naphthol derivatives with acyl halides to yield the corresponding ketones in high yields.⁵ The hydroxynaphthyl ketones (4) can also be obtained by $\text{Sc}(\text{OTf})_3$ -catalysed Fries rearrangement of acyloxy naphthalenes (Scheme 11.2).⁶



Scheme 11.2

$\text{Sc}(\text{OTf})_3$ catalyses Friedel–Crafts benzylation and allylation reactions with arenes using benzyl alcohol (6) and allylic alcohols (8) as electrophiles to give diarylmethanes (7) and allylarenes (9), respectively (Scheme 11.3).⁵ Water,

11: Scandium(III) and yttrium(III)



Scheme 11.3

produced during the reaction, does not spoil catalytic activity of Sc(OTf)_3 and a catalytic amount (10 mol %) of Sc(OTf)_3 is enough for the reaction. In contrast to conventional Friedel–Crafts benzylation and allylation using organic halides as electrophiles which produces troublesome hydrogen halides as by-products, this process only produces water. Considering the advantages of Sc(OTf)_3 in the Friedel–Crafts reaction, it may solve some severe environmental problems induced by conventional Lewis acid-promoted reactions in industry.

Protocol 1.

Preparation of *p*-methoxyacetophenone⁴

Equipment

- Three-necked round-bottomed flask (50 mL)
- Magnetic stirrer
- Thermostatted oil bath
- Reflux condenser
- Three-way stopcock
- Stopcocks
- Thermometer (100°C)
- Drying tube (calcium chloride)
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- Erlenmeyer flask (200 mL)
- Teflon-coated magnetic stirrer bar (2.0 × 0.7 cm)

Materials

- Distilled (CaH₂ nitromethane (FW 61.04), 10 mL
- Anisole (FW 108.14), 1.08 g, 10.0 mmol
- Distilled acetic anhydride (FW 102.09), 1.02 g, 10.0 mmol
- Sc(OTf)_3 (FW 492.16), 0.98 g, 2.0 mmol
- Diethyl ether for extraction, 60 mL
- Anhydrous magnesium sulfate

flammable
irritant, hygroscopic
corrosive, lachrymator
irritant, hygroscopic
flammable, toxic
hygroscopic

1. Place Sc(OTf)_3 (0.98 g, 2.0 mmol) and a magnetic stirrer bar in a three-necked round-bottomed flask (50 mL) connected with two stopcocks and a three-way stopcock.
2. Heat a flask in an oil bath at 180°C under vacuum (1 mm Hg) for 1 h.
3. When the flask has cooled to room temperature, connect a reflux condenser, a thermometer and a drying tube.

Protocol 1. Continued

4. Add nitromethane (10 mL) to the flask and stir for 10 min, and then add anisole (1.08 g, 10.0 mmol) and acetic anhydride (1.02 g, 10.0 mmol).
5. Heat the solution to 50°C in an oil bath equipped with a temperature controller (bath temperature, 70°C).
6. Keep the temperature at 50°C with stirring and monitor the disappearance of anisole by TLC (visualized with UV). It takes 4 h to complete the reaction.
7. Add water (30 mL) to the solution. Separate the organic layer with the aid of separating funnel (100 mL). Extract the aqueous layer with three portions of diethyl ether (20 mL). Wash the combined organic layers with brine (30 mL). Transfer the solution to an Erlenmeyer flask and dry the solution over anhydrous magnesium sulfate.
8. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator to leave a pale-yellow residue. Distill the residue by Kugelrohr at 0.4 mmHg to collect 1.22 g (7.4 mmol, 74%) of the product (oven temperature, 110°C). The compound was characterized by ¹H NMR and elemental analysis. The product comprises only *p*-acetoxyanisole; no *o*- and *m*-isomers were detected by capillary GC analysis (capillary column coated with (5%-phenyl)methylpolysiloxane).

Protocol 2.**Reaction of benzylalcohol with benzene⁷***Equipment*

- Three-necked round-bottomed flask (200 mL)
- Magnetic stirrer
- Thermostatted oil bath
- Reflux condenser
- Three-way stopcock
- Drying tube (calcium chloride)
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3).
- Erlenmeyer flask (200 mL)
- Teflon-coated magnetic stirrer bar (2.0 × 0.7 cm)

Materials

- Benzene (FW 78.11), 100 mL, 1.11 mol
- Distilled benzyl alcohol (FW 108.14) 2.16 g, 20.0 mmol
- Sc(OTf)₃^a (FW 492.16), 0.98 g, 2.0 mmol
- Diethyl ether for extraction, 50 mL
- Anhydrous magnesium sulfate

cancer-suspect agent, flammable
irritant, hygroscopic
irritant, hygroscopic
flammable, irritant
hygroscopic

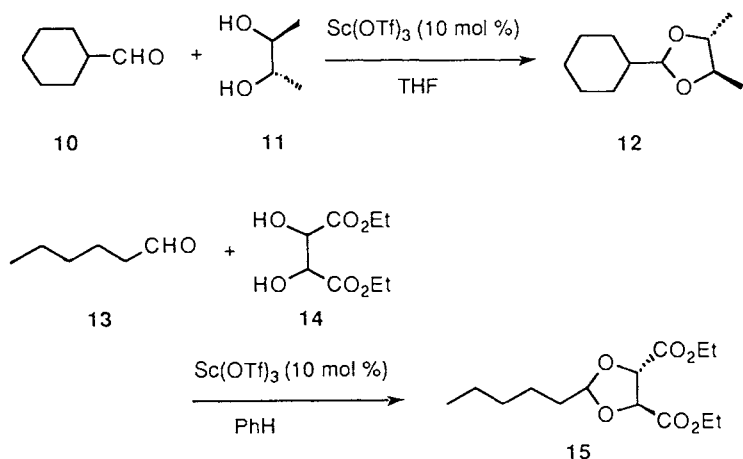
1. Assemble a three-neck round-bottomed flask, a magnetic stirrer bar, a reflux condenser, and a drying tube.
2. Add Sc(OTf)₃ (0.98 g, 2.0 mmol), benzene (100 mL) and benzyl alcohol (2.16 g, 20.0 mmol) to the flask and heat the solution at refluxing temperature with stirring.
3. After stirring for 8 h under reflux, cool the flask to room temperature.

4. Add water (50 mL) to the solution. Separate the organic layer with the aid of separating funnel (300 mL). Extract the aqueous layer with diethyl ether (50 ml). Wash the combined organic layers with brine (30 ml). Transfer the organic layer to an Erlenmeyer flask and dry the solution over anhydrous magnesium sulfate.
5. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator. Distil the residue at 3.0 mmHg to collect 2.7 g (16.2 mmol, 81%) of the product boiling at 97°C. The compound was characterized by GC/MS, ^1H NMR and elemental analysis.
6. Concentrate the aqueous layer with a rotary evaporator and heat the crystalline residue $[\text{Sc}(\text{OTf})_3]$ at 180°C under vacuum for 20 h. The recovered $\text{Sc}(\text{OTf})_3$ can be used in the next Friedel–Crafts benzylation and allylation reactions.

^a Commercial $\text{Sc}(\text{OTf})_3$ retains moisture. Dry $\text{Sc}(\text{OTf})_3$ by heating in a flask using an oil bath at 180°C under vacuum (1 mm Hg) for more than 1 h (see Protocol 1).

2.2 Acetalization

Chiral acetals are particularly important precursors for the synthesis of enantiomerically pure compounds.⁸ Chiral dioxane and dioxolane are prepared by the direct reaction of aldehydes or ketones with chiral 1,3-diols and 1,2-diols in the presence of a catalytic amount of $\text{Sc}(\text{OTf})_3$ at room temperature (Scheme 11.4).⁹ Chiral 2,4-pentandiol, 2,3-butandiol, and diethyl tartarate are used as chiral diols. Considering that the chiral acetals derived from diethyl tartarate have only been prepared by trans-acetalization under acidic conditions,^{10,11} this method is convenient and practical. It is worth noting that removal of water which results from the acetalization is not necessary.



Scheme 11.4

Protocol 3.**Scandium(III) triflate catalysed acetalization of aldehyde⁹***Equipment*

- Three-necked round-bottomed flask (100 mL)
- Magnetic stirrer
- Three-way stopcock
- Stopcocks
- Drying tube (calcium chloride)
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- Erlenmeyer flask (200 mL)
- Teflon-coated magnetic stirrer bar (2.0 × 0.7 cm)

Materials

- Dry distilled benzene (P₂O₅) (FW 78.11), 30 mL **cancer-suspect agent, flammable**
- Distilled hexanal (FW 100.16), 1.00 g, 10.0 mmol **flammable, irritant**
- L-(+)-Diethyl tartarate (FW 206.19), 2.50 g, 12.1 mmol
- Sc(OTf)₃^a (FW 492.16), 0.98 g, 2.0 mmol **irritant, hygroscopic**
- Diethyl ether for extraction, 50 mL **flammable, irritant**
- Anhydrous magnesium sulfate **hygroscopic**

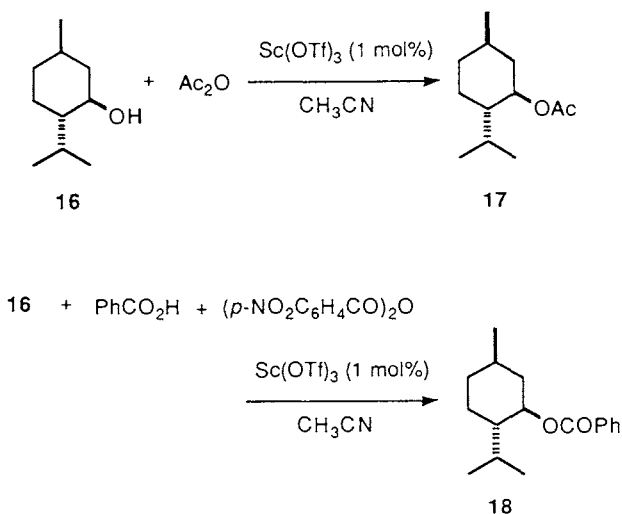
1. Assemble a three-neck round-bottomed flask (100 mL), a magnetic stirrer bar, a stopcock, a three-way stopcock and a drying tube.
2. Add Sc(OTf)₃ (0.98 g, 2.0 mmol) and benzene (15 mL) into a flask and stir the suspension for 10 min at room temperature.
3. Add (1*R*, 2*R*)-(+)-diethyl tartarate (2.50 g, 12.1 mmol) and stir the mixture at room temperature for 5 min until the mixture becomes homogeneous. Then, add hexanal (1.0 g, 10.0 mmol) and stir the solution at room temperature.
4. Monitor the production of the acetal by GC. After stirring for 48 h, add water (50 mL) to the solution. Separate the organic layer with the aid of a separating funnel (300 mL). Extract the aqueous layer with diethyl ether (50 mL). Wash the combined organic layers with brine (30 mL). Transfer the organic layer to a Erlenmeyer flask (200 mL) and then dry them over anhydrous magnesium sulfate.
5. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator. Distil the residue at 0.3 mmHg to collect 1.98 g (6.9 mmol, 69%) of the product boiling at 97–100°C. The compound was characterized by GC/MS, ¹H NMR and elemental analysis.

^aCommercial Sc(OTf)₃ retains moisture. Dry Sc(OTf)₃ by heating in a flask using oil bath at 180°C under vacuum (1 mm HG) for more than 1 h (see Protocol 1).

2.3 Acetylation and esterification of alcohol

Acylation of alcohol is often required in organic synthesis, e.g. protection of a hydroxy group. The base-catalysed acylation by acid anhydride or acyl chloride is usually performed with a tertiary amine such as pyridine and its derivatives.¹² Acid-catalysed acylation of alcohols with acid anhydride is an

alternative acylation method which requires mild conditions. $\text{Sc}(\text{OTf})_3$ is an efficient catalyst for acylation of alcohols by acid anhydrides (Scheme 11.5).^{13,14} It also effectively catalyses the esterification of alcohols with carboxylic acid in the presence of *p*-nitrobenzoic anhydride.



Scheme 11.5

Protocol 4.

Preparation of L-acetyl menthol (17)^{13,14}

Equipment

- Three-necked round-bottomed flask (100 mL)
- Magnetic stirrer
- Three-way stopcock
- Stopcocks
- Drying tube (calcium chloride)
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- All-glass syringe with a needle-lock luer (volume appropriate for quantity of solution to be transferred)
- Erlenmeyer flask (200 mL)
- Teflon-coated magnetic stirrer bar (2.0×0.7 cm)
- Glass column (2.5 cm \times 50 cm) packed with silica gel

Materials

- Dry distilled acetonitrile (CaH_2) (FW 41.05), 30 mL
- Distilled acetic anhydride (FW 102.09), 1.0 mL, 10.6 mmol
- L-(-)-Menthol (FW 156.27), 1.56 g, 10.0 mmol
- $\text{Sc}(\text{OTf})_3^a$ (FW 492.16), 0.05 g, 0.1 mmol
- Diethyl ether for extraction, 150 mL
- n-Hexane for column chromatography
- Ethyl acetate for column chromatography
- Anhydrous magnesium sulfate

flammable, toxic
corrosive, lachrymator
irritant
irritant, hygroscopic
flammable, irritant
flammable, irritant
irritant
hygroscopic

Protocol 4. Continued

1. Assemble a three-necked round-bottomed flask (100 mL), drying tube, a three-way stopcock, and a stopcock.
2. Add L-menthol (1.56 g, 10.0 mmol), acetic anhydride (1.53 g, 15.0 mmol), and acetonitrile (29 mL) to the flask and stir the solution for 10 min at room temperature.
3. Add dropwise a acetonitrile solution (1.0 mL) of $\text{Sc}(\text{OTf})_3$ (49.6 mg, 0.10 mmol) with the aid of a syringe (2.0 ml) through the three-way stopcock.
4. After stirring for 1 h, add a saturated aqueous sodium hydrogen carbonate solution (50 mL) to the solution. Transfer the mixture to a separating funnel (300 mL) and extract the aqueous layer with three portions of diethyl ether (50 mL) and separate the organic layer. Wash the combined organic layers with brine (30 mL), then dry over anhydrous magnesium sulfate. Capillary GC analysis (capillary column coated with 100% dimethylpolysiloxane) of the organic solution shows that L-menthol is converted into the acylated product quantitatively.
5. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator. Prepare a column (2.5 cm \times 50 cm) for chromatography using silica gel (Merck Silica Gel 60) and *n*-hexane as the eluant. Dissolve the residue in a minimum volume (2 mL) of *n*-hexane and elute the column with *n*-hexane–ethyl acetate (5:1). Evaporate the eluate on a rotary evaporator to yield a colourless residue (2.00 g, 10.0 mmol, 100%). GC/MS and ^1H NMR spectra of the products display the suitable structure.

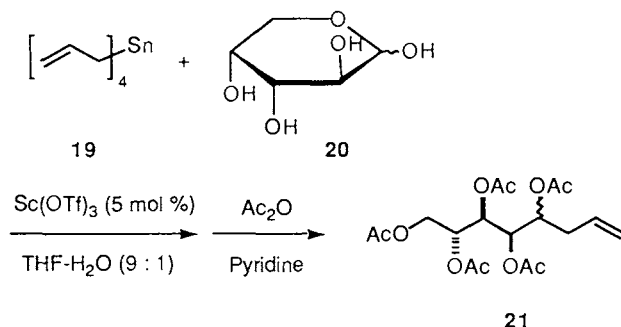
^a Commercial $\text{Sc}(\text{OTf})_3$ retains moisture. Dry $\text{Sc}(\text{OTf})_3$ by heating in a flask using an oil bath at 180°C under vacuum (1 mm Hg) for more than 1 h (see Protocol 1).

2.4 Allylation of carbonyl compounds with tetra-allyltin catalysed by scandium(III) triflate

Allylation of carbonyl compounds has been recognized as an important process to produce synthetically useful homoallylic alcohols.¹⁵ Tetra-allyltin is a good allyl transfer agent and itself reacts with ketones and aldehydes in the presence of an acid catalyst under mild conditions.¹⁶ Water-tolerant $\text{Sc}(\text{OTf})_3$ is an efficient catalyst for the allylation of carbonyl compounds with tetra-allyltin in aqueous media.¹⁷ The allylation reaction of ketones and aldehydes proceeds smoothly in the presence of catalytic amounts of $\text{Sc}(\text{OTf})_3$ to produce homoallylic alcohols in good yields under mild conditions. For example, monosaccharides such as D-arabinose (**20**) is easily allylated in aqueous media (Scheme 11.6).

The allylated adducts of saccharides are intermediates for the synthesis of higher saccharides.

11: Scandium(III) and yttrium(III)



Scheme 11.6

Protocol 5.

Scandium(III) triflate catalysed allylation of D-arabinose¹⁷

Equipment

- Three-necked round-bottomed flask (200 mL)
- Single-necked round-bottomed flask (300 mL)
- Magnetic stirrer
- Three-way stopcock
- Drying tube (calcium chloride)
- Pressure equalizing dropping funnel (50 mL)
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- Erlenmeyer flask (200 mL)
- Teflon-coated magnetic stirrer bar (2.0 × 0.7 cm)
- Glass column (2.5 cm × 50 cm) packed with silica gel

Materials

- Dry distilled acetonitrile (CaH₂) (FW 41.05), 90 mL
 - Distilled water (FW 18.02), 10 mL
 - D-(+)-Arabinose (FW 151.12), 1.51 g, 10.0 mmol
 - Tetra-allyltin^a (FW 282.98), 1.41 g, 5.0 mmol
 - Sc(OTf)₃^b (FW 492.16), 0.25 g, 0.5 mmol
 - Distilled pyridine (CaH₂) (FW 79.10), 25 mL
 - Acetic anhydride (FW 102.09), 10 mL, 0.13 mmol
 - Diethyl ether for extraction, 150 mL
 - Ethyl acetate for column chromatography
 - Anhydrous magnesium sulfate
- flammable, toxic**

toxic
irritant, hygroscopic
flammable, toxic
corrosive, lachrymator
flammable, irritant
flammable, irritant
hygroscopic

1. Assemble a three-necked round-bottomed flask (200 mL), drying tube, a three-way stopcock, and a dropping funnel.
2. Add Sc(OTf)₃ (0.25 g, 0.50 mmol) and acetonitrile (40 mL)–water (10 mL) solution of D-arabinose (1.51 g, 10.0 mmol) to the flask.
3. Add acetonitrile (50 mL) solution of tetra-allyltin (1.41 g, 5.0 mmol) dropwise from a dropping funnel at room temperature and stir the solution at the same temperature for 60 h.
4. Transfer the solution into single-necked round-bottomed flask (300 mL) and remove the solvents on a rotary evaporator to leave a white residue.

Protocol 5. Continued

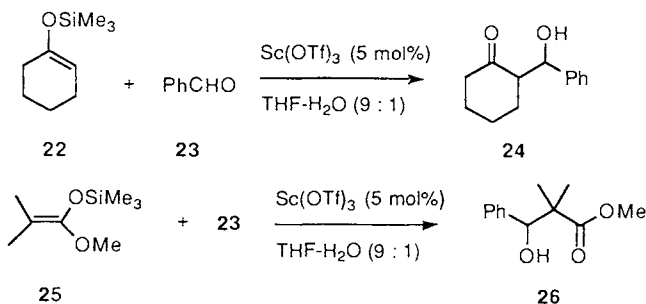
5. Cool the flask with a ice-bath and add pyridine (25 mL) and Ac_2O (10 mL, 0.13 mol) with stirring. After 30 min warm the flask to room temperature and stir for 5 h.
6. Transfer the solution into single-necked round-bottomed-flask (200 mL) and remove 70–80% volume of pyridine and excess of Ac_2O using a rotary evaporator under reduced pressure (10 mm Hg, temp 70°C). Treat the residue with cold 1M HCl (100 mL) and transfer the mixture to a separating funnel (300 mL). Extract the water layer with three portions of diethyl ether (50 ml) and wash the combined organic layer with a saturated aqueous solution of NaHCO_3 (50 mL). Transfer the organic layer to an erlenmeyer flask and dry the solution over anhydrous magnesium sulfate.
7. Remove the solvent on a rotary evaporator to leave a pale-yellow viscous residue. Prepare a column (2.5 cm \times 50 cm) for chromatography using silica gel (Merck Silica Gel 60) and hexane as the eluant. Dissolve the residue in a minimum volume (5 mL) of chloroform and elute the column with ethyl acetate. Evaporate the ethyl acetate eluate on a rotary evaporator to yield a colourless residue (2.78 g, 6.9 mmol, 69%). Capillary GC analysis (Capillary column coated with 100% dimethypolysiloxane) of the residue shows that the product is a mixture of diastereomers (*anti:syn* 5 65:35). ^1H NMR spectra of the products display suitable structures.

^a Tetra-allyltin is prepared by the reaction of allylmagnesium chloride with tin(IV) chloride.^{18,19}

^b Commercial $\text{Sc}(\text{OTf})_3$ retains moisture. Dry $\text{Sc}(\text{OTf})_3$ by heating in a flask using oil bath at 180°C under vacuum (1 mm Hg) for more than 1 h (see Protocol 1).

2.5 Scandium(III) triflate-catalysed Mukaiyama aldol reaction

Although a wide variety of Lewis acid catalysts has been proposed for the Mukaiyama aldol reaction,²⁰ $\text{Sc}(\text{OTf})_3$ is a promising alternative catalyst for the reaction from the viewpoint that it can be used in aqueous media and recovered and reused (Scheme 11.7).²¹ Most Lewis acids usually decompose



Scheme 11.7

in the presence of even a small amount of water but $\text{Sc}(\text{OTf})_3$ works as a catalyst in aqueous media. A catalytic amount of $\text{Sc}(\text{OTf})_3$ (5 mol%) effectively catalyses the reaction of silyl enol ethers with aldehyde or acetals to yield the corresponding aldol adducts in H_2O –THF (1:9) solvent.

Protocol 6.

$\text{Sc}(\text{OTf})_3$ -catalysed reaction of 1-trimethylsiloxy-1-cyclohexene (22) with benzaldehyde (23) in aqueous media²¹

Equipment

- Three-necked round-bottomed flask (100 mL)
- Magnetic stirrer
- Three-way stopcock
- Stopcocks
- Septum
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- All-glass syringe with a needle-lock luer (volume appropriate for quantity of solution to be transferred)
- Erlenmeyer flask (100 mL)
- Teflon-coated magnetic stirrer bar (2.0×0.7 cm)
- Glass column (2.5 cm \times 50 cm) packed with silica gel

Materials

- | | |
|---|-----------------------------|
| • Distilled dry tetrahydrofuran (THF) (FW 80.17), 25 mL | flammable, irritant |
| • 1-Trimethylsiloxy-cyclohexene ^a (FW 170.33), 1.70 g, 10.0 mmol | moisture sensitive irritant |
| • Distilled benzaldehyde (FW 106.12), 1.06 g, 10.0 mmol | cancer suspect |
| • $\text{Sc}(\text{OTf})_3$ (FW 492.16), ^b 0.25 g, 0.5 mmol | irritant, hygroscopic |
| • Diethyl ether for extraction, 150 mL | flammable, irritant |
| • <i>n</i> -Hexane for column chromatography | flammable, irritant |
| • Ethyl acetate for column chromatography | flammable, irritant |
| • Anhydrous magnesium sulfate | hygroscopic |

1. Assemble a three-necked round-bottomed flask (100 mL), a three-way stopcock and stopcocks under argon while the apparatus is still hot.
2. Assemble a syringe (20 mL) and needle (30 cm) while hot and allow the assembled syringe to cool to room temperature in a desiccator. Flush the syringe with argon.
3. When the flask has cooled to room temperature, add $\text{Sc}(\text{OTf})_3$ (0.246 g, 0.50 mmol). Then, add THF (8.0 mL) and H_2O (2.0 mL) to the flask using the syringe through the septum on the three-way stopcock.
4. Add THF (10.0 mL) solution of benzaldehyde (1.06 g, 10.0 mmol) and 1-trimethylsiloxy-1-cyclohexene (1.72 g, 10.0 mmol) using a syringe (20 mL) through the septum on the three-way stopcock at room temperature. Stir the mixture and monitor the reaction by a silica gel TLC (Merck Silica Gel 60 PF₂₅₄).
5. After 7 days, add water (50 mL) to the solution. Transfer the mixture to a separating funnel (300 mL) and extract the aqueous layer with three portions of diethyl ether (50 mL). Wash the combined organic layers with brine (30 mL), then dry over anhydrous magnesium sulfate. Capillary GC analysis

Protocol 6. Continued

(capillary column coated with (5%-phenyl)methylpolysiloxane) of the organic solution shows the presence of a diastereomeric mixture of the product (*syn:anti* = 75:25).

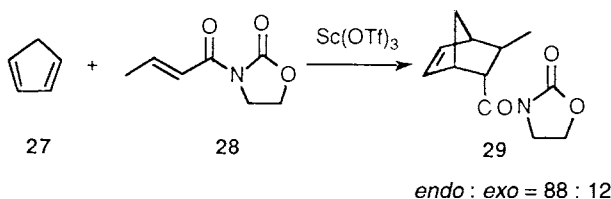
6. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator. Prepare a column (2.5 cm × 50 cm) for chromatography using silica gel (Merck Silica Gel 60) and *n*-hexane as the eluant. Dissolve the residue in a minimum volume (2 mL) of *n*-hexane and elute the column with *n*-hexane–ethyl acetate (5:1). Evaporate the eluate on a rotary evaporator to yield a white solid (1.28 g, 6.3 mmol, 63%). GC/MS and ¹H NMR spectra of the products display the suitable structure.

^a Commercially available.

^b Commercial Sc(OTf)₃ retains moisture. Dry Sc(OTf)₃ by heating in a flask using an oil bath at 180°C under vacuum (1 mm Hg) for more than 1 h (see Protocol 1).

2.6 Diels–Alder reaction

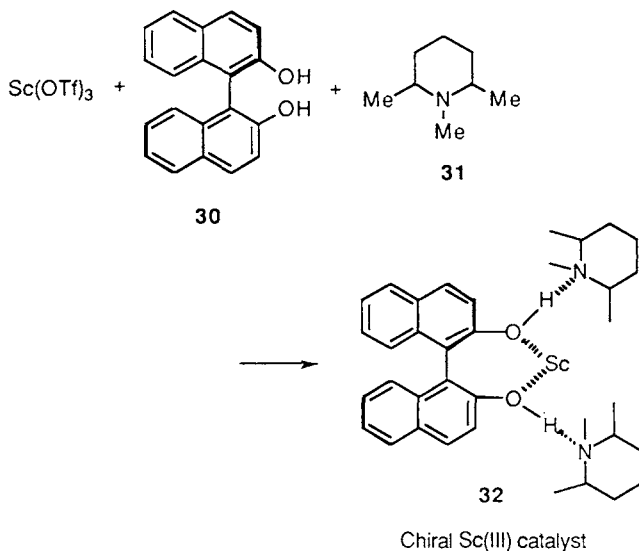
Although Lewis acid-catalysed Diels–Alder reactions have been shown to proceed under mild conditions, they are often accompanied by a diene polymerization.²² They frequently need excess amounts of the catalyst in the carbonyl-containing dienophiles. Sc(OTf)₃ is an efficient catalyst for Diels–Alder reactions of carbonyl-containing dienophiles with cyclopentadiene.²³ Thus, 3-crotonoyl-1,3-oxazolidin-2-one (**28**) reacts with cyclopentadiene (**27**) to give a mixture of corresponding adduct (**29**), in the ratio of *endo:exo* = 87:13, in a good yield (Scheme 11.8).



Scheme 11.8

Asymmetric version of this reaction can be achieved using chiral binaphthol (**30**), *cis*-1,2,3-trimethylpiperidine (**31**), and molecular sieve 4A together with Sc(OTf)₃ (Scheme 11.9).²⁴ The *endo* adduct (*endo:exo* = 86:14–89:11) of the reaction of **28** with **27** is obtained in up to 96% ee. It must be noted that the enantioselectivity of the reaction decreases in accordance with ageing time and temperature of the preparation of the chiral catalyst (**32**).²⁵ 3-Acetyl-1,3-oxazolidin-2-one is found to be effective for preventing the ageing effect.

11: Scandium(III) and yttrium(III)



Scheme 11.9

Protocol 7.

Asymmetric Diels–Alder reaction of cyclopentadiene (27) with 3-crotonoyl-1,3-oxazolidin-2-one (28)²⁴

Equipment

- Three-necked round-bottomed flask (100 mL)
- Magnetic stirrer
- Three-way stopcock
- Stopcocks
- Septum
- All-glass syringe with a needle-lock luer (volume appropriate for quantity of solution to be transferred)
- Dry ice–methanol bath
- Separating funnel (300 mL)
- Sintered-glass filter funnel (porosity 3)
- Erlenmeyer flask (100 mL)
- Glass column (2.5 cm × 50 cm) packed with silica gel
- Vacuum/inert gas source and inlet

Materials

- Distilled (P₂O₅) dichloromethane (FW 84.93), 25 mL toxic, irritant
- 3-Crotonoyl-1,3-oxazolidin-2-one^a (FW 155.15), 1.55 g, 10 mmol
- Freshly distilled cyclopentadiene^b (FW 66.10), 1.98 g, 30.0 mmol flammable
- Powder molecular sieve 4A (activated)^c 1.0 g
- (*R*)-Binaphthol^d (FW 286.33), 0.34 g, 1.18 mmol irritant
- *cis*-1,2,3-Trimethylpiperidine^e (FW 127.22), 0.31 g, 2.4 mmol
- Sc(OTf)₃ (FW 492.16), 0.49 g, 1.0 mmol irritant, hygroscopic
- Diethyl ether for extraction, 150 mL flammable, toxic
- *n*-Hexane for column chromatography flammable, irritant
- Ethyl acetate for column chromatography flammable, irritant
- Anhydrous magnesium sulfate hygroscopic

Protocol 7. Continued

1. Place $\text{Sc}(\text{OTf})_3$ (0.49 g, 1.0 mmol) and powder molecular sieves 4A (activated) (1.0 g) in a three-necked round-bottomed flask (100 mL) connected with a three-way stopcock and stopcocks.
2. Heat a flask by oil bath at 190°C under vacuum (1 mm Hg) for 5 h.
3. Flush the flask with argon through several vacuum cycles while the apparatus is still hot.
4. After cooling the flask to room temperature, add (*R*)-binaphthol (0.34 g, 1.2 mmol) to the flask.
5. Assemble a syringe (20 mL) and needle (30 cm) while hot and allow the assembled syringe to cool to room temperature in a desiccator. Flush the syringe with argon.
6. Add dichloromethane (10 mL) using the syringe through the septum on the three-way cock and cool the flask with dry ice-ethanol bath to -78°C under a slight pressure of argon.
7. Add *cis*-1,2,3-trimethylpiperidine (0.31 g, 2.4 mmol) using a syringe (1 mL) through the septum on the three-way stopcock into the flask at -78°C . Stir the mixture at this temperature for 30 min.
8. Add dichloromethane (15.0 mL) solution of 3-crotonoyl-1,3-oxazolidin-2-one (1.55 g, 10 mmol) and freshly distilled cyclopentadiene (1.98 g, 30 mmol) using a syringe (10 mL) through the septum on the three-way stopcock.
9. Stir the mixture at -78°C for 4 h and then warm it slowly to 0°C over a period of 10 h. Monitor the reaction by a silica gel TLC (Merck Silica Gel 60 PF₂₅₄).
10. Add water (50 mL) to the solution and transfer the mixture to a separating funnel (300 mL) and separate the organic layer. Extract the aqueous layer with two portions of diethyl ether (50 mL). Wash the combined organic layers with brine (30 mL), then dry them over anhydrous magnesium sulfate. Capillary GC analysis (capillary column coated with (5%-phenyl)methylpolysiloxane) of the organic solution shows the presence of a diastereomeric mixture of the product (*endo:exo* = 88:12).
11. Filter the dried solution through a sintered-glass filter funnel and remove the solvent on a rotary evaporator. Prepare a column (2.5 cm \times 50 cm) for chromatography using silica gel (Merck Silica Gel 60) and *n*-hexane as the eluant. Dissolve the residue in a minimum volume (2 mL) of *n*-hexane and elute the column with *n*-hexane-ethyl acetate (10:1). Evaporate the eluate on a rotary evaporator to yield a colourless residue (1.40 g, 7.0 mmol, 70%). GC/MS and ^1H NMR spectra of the products display the suitable structure.

12. The enantiomeric excess of the *endo* adduct is determined to be 86% ee by HPLC analysis (Daicel Ciralpak AD; *n*-hexane-*i*-PrOH, 80:20; 0.5 ml/min).

^a 3-Crotonoyl-1,3-oxazolidin-2-one is prepared by the reaction of crotonoyl chloride with 2-oxazolidinone in THF in the presence of *n*-butyl-lithium.²⁶

^b Cyclopentadiene is freshly distilled by thermolysis of commercial dicyclopentadiene at 180°C just prior to use.

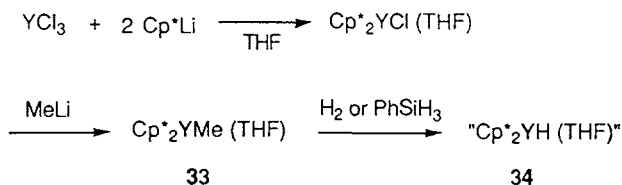
^c Obtained from Aldrich.

^d Commercially available.

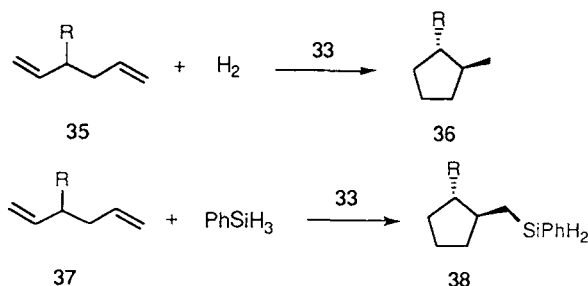
^e *cis*-1,2,3-trimethylpiperidine is prepared by *N*-methylation of commercial *cis*-1,3-dimethylpiperidine by formaldehyde and formic acid.²⁷

3. Cyclopentadienyl yttrium hydride

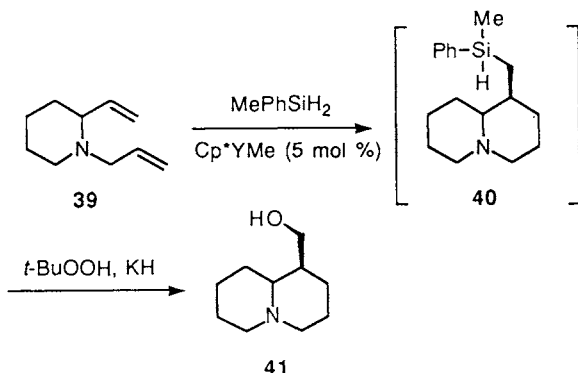
Since inorganic yttrium (III) compounds are less Lewis acidic than the corresponding scandium(III) compounds, they are seldom used as Lewis acid catalysts. Cyclopentadienyl derivatives of yttrium(III) hydrides have been studied in organic synthesis and found to catalyse hydrosilylation of alkenes²⁸ and cyclization of dienes.²⁹ The high regioselectivities and stereoselectivities of these reactions are worthy of note. Cyclopentadienyl yttrium hydride [$(\eta^5\text{-C}_5\text{Me}_5)_2\text{YH}(\text{THF})$] (**34**) is conveniently prepared *in situ* in the reaction of ciscyclopentadienylalkyl yttrium [$(\eta^5\text{-C}_5\text{Me}_5)_2\text{YCH}_3(\text{THF})$] (**33**) with hydrogen or phenylsilane and can be used for the catalytic reactions (Scheme 11.10). Thus, $(\eta^5\text{-C}_5\text{Me}_5)_2\text{YCH}_3(\text{THF})$ catalyses the cyclization/hydrosilylation of dienes with phenylsilane smoothly to afford the cyclized organosilane (**38**) (Scheme 11.11).³⁰ Cyclization reaction of dienes has been applied to the synthesis of epilupine, one of the lupin alkaloids (**41**) (Scheme 11.12).³¹



Scheme 11.10



Scheme 11.11



Scheme 11.12

References

1. Cotton, F. A.; Wilkinson, G. *Advanced inorganic chemistry* 5th edn. Wiley, 1988, New York.
2. Kobayashi, S. *Synlett* **1994**, 689.
3. *Synthetic methods of organometallic and inorganic chemistry*, Vol 6; Lanthanides and Actinides (ed. F. T. Edelman). Thieme, New York, **1997**, pp 36–37.
4. Kawada, A.; Mitamura, S.; Kobayashi, S. *Synlett* **1994**, 545.
5. Kobayashi, S.; Moriwaki, M.; Hachiya, I. *Synlett* **1995**, 1153.
6. Kobayashi, S.; Moriwaki, M.; Hachiya, I. *J. Chem. Soc. Chem. Commun.* **1995**, 1527.
7. Tsuchimoto, T.; Tobita, K.; Hiyama, T.; Fukuzawa, S. *Synlett* **1996**, 557.
8. Alexakis, A.; Mangeney, P. *Tetrahedron Asymmetry* **1990**, 1, 477.
9. Fukuzawa, S.; Tsuchimoto, T.; Hotaka, T.; Hiyama, T. *Synlett* **1995**, 1077.
10. Fukutani, Y.; Maruoka, K.; Yamamoto, H. *Tetrahedron Lett.* **1984**, 25, 5911.
11. Fujiwara, J.; Fukutani, Y.; Hasegawa, M.; Maruoka, K.; Yamamoto, H. *J. Am. Chem. Soc.* **1984**, 106, 5004.
12. Hoefle, G.; Steilich, V.; Vorbruggen, H. *Angew. Chem., Int. Ed. Engl.* **1978**, 17, 569.
13. Ishihara, K.; Kubota, M.; Kurihara, H.; Yamamoto, H. *J. Am. Chem. Soc.* **1995**, 117, 4413.
14. Ishihara, K.; Kubota, M.; Kurihara, H.; Yamamoto, H. *J. Org. Chem.* **1996**, 61, 4560.
15. Roush, W. R. *Comprehensive organic synthesis*, Vol 2 (ed. B. M. Trost). Pergamon Press, Oxford, **1991**, pp 1–53.
16. Yanagisawa, A.; Inoue, H.; Morodome, M.; Yamamoto, H. *J. Am. Chem. Soc.* **1993**, 115, 10356.
17. Hachiya, I.; Kobayashi, S. *J. Org. Chem.* **1993**, 58, 6958.
18. Fishwick, M.; Wallbridge, M. G. H. *J. Organomet. Chem.* **1970**, 25, 69.
19. Daude, G.; Pereyre, M. *J. Organomet. Chem.* **1980**, 190, 43.
20. Gennari, C. *Comprehensive organic synthesis*, Vol 2 (ed. B. M. Trost). Pergamon Press, Oxford, **1991**, pp 629–660.

21. Kobayashi, S.; Hachiya, I.; Ishitani, H.; Araki, M. *Synlett* **1993**, 472.
22. Carruthers, W. *Cycloaddition reactions in organic synthesis*, Pergamon Press, Oxford, 1990.
23. Kobayashi, S.; Hachiya, I.; Araki, M.; Ishitani, H. *Tetrahedron Lett.* **1993**, 34, 3755.
24. Kobayashi, S.; Araki, M.; Hachiya, I. *J. Org. Chem.* **1994**, 59, 3758.
25. Kobayashi, S.; Ishitani, H.; Araki, M.; Hachiya, I. *Tetrahedron Lett.* **1994**, 35, 6325.
26. Evans, D. A.; Chapman, K. T.; Bisaha, J. *J. Am. Chem. Soc.* **1988**, 110, 1238.
27. Icke, R. N.; Wisegarar, B. B.; Alles, G. A. *Organic synthesis*. Wiley, New York, **1955**, Collect. Vol 3, p 723.
28. Molander, G. A.; Julius, M. *J. Org. Chem.* **1992**, 57, 6347.
29. Molander, G. A.; Hoberg, J. O. *J. Am. Chem. Soc.* **1992**, 114, 3123.
30. Molander, G. A.; Nichols, P. J. *J. Am. Chem. Soc.* **1995**, 117, 4415.
31. Molander, G. A.; Nichols, P. J. *J. Org. Chem.* **1996**, 61, 6040.

Lanthanide(III) Reagents

TAKESHI NAKAI and KATSUHIKO TOMOOKA

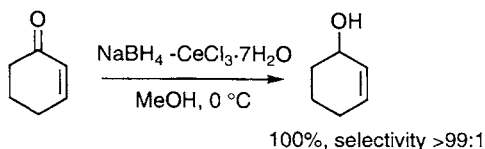
1 Introduction

The lanthanide group consists of the 15 elements from lanthanum to lutetium in the periodic table. In the past two decades the synthetic uses of a variety of lanthanide complexes, in terms of the kind of central lanthanides, the oxidation state (+2–+4), and the kind of ligands, have been developed.¹ This chapter focuses primarily on the four classes of the most popular trivalent lanthanides, namely, cerium(III) complexes (CeL_3), europium(III) complexes (EuL_3), ytterbium complexes (YbL_3), and lanthanum(III) complexes (LaL_3). Note that many synthetic applications have been found not only of divalent samarium reagents such as SmI_2^{2-4} and tetravalent cerium reagents such as cerium ammonium nitrate (CAN)⁵ but also of lanthanide metals such as Ce, Sm, and Yb, which are, however, beyond the scope of this chapter.

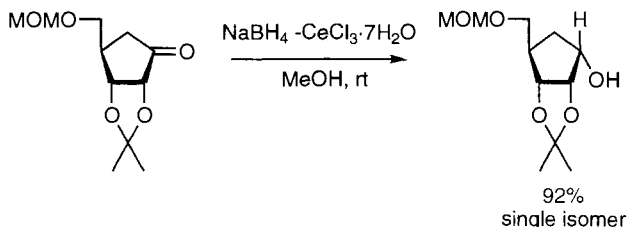
The trivalent state is the most common oxidation state for lanthanides. Generally speaking, lanthanide(III) complexes (except for the alkoxides) are hard Lewis acids and hence they have a strong affinity toward hard bases such as oxygen donor ligands. This strong oxophilicity is one of the most important characteristics of lanthanide(III) reagents and provides the principal basis for their unique synthetic applications.

2. Cerium(III) reagents

Cerium(III) salts, particularly CeCl_3 , have found many synthetic applications as inexpensive, unique reagents (stoichiometric use). Of the most widespread use are the NaBH_4 – CeCl_3 system that allows the regioselective reductions of α,β -enones (Scheme 12.1)^{6,7} and the stereoselective reduction of ketones

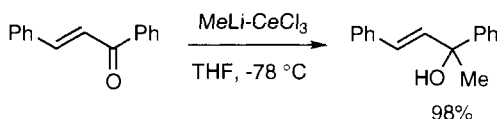


Scheme 12.1



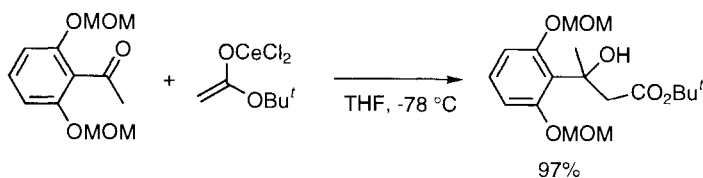
Scheme 12.2

(Scheme 12.2),⁸ and the organocerium reagents (RCeCl_2), generated *in situ* from organolithiums or Grignard reagents and CeCl_3 , which can smoothly undergo the normal 1,2-addition reactions to carbonyl compounds including α,β -enones (Scheme 12.3).⁹ Particularly notable is that the use of CeCl_3 as an additive in these reactions significantly suppresses unfavourable side reactions (enolization, reduction, conjugate addition, etc.) which are often encountered with the use of organolithiums or Grignard reagents alone. In addition, a higher stereoselectivity is often observed by virtue of the chelating ability of the cerium reagents involved.



Scheme 12.3

In a similar manner, treatment of lithium enolates with CeCl_3 generates the cerium enolates which undergo aldol reactions usually in a higher yield than with lithium enolates alone,^{10,11} particularly when sterically hindered and/or easily enolizable carbonyl partners are used (Scheme 12.4).¹¹ The stereoselectivity of Ce -enolate reactions is essentially the same as that of Li -enolate reactions, since both reactions proceed through a cyclic transition state.



Scheme 12.4

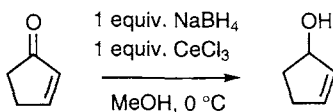
Furthermore, cerium halides are also useful as promoters for the reductive coupling of α -haloketones with carbonyl compounds. The use of CeI_3 affords

the α,β -unsaturated carbonyl compounds, whereas the $\text{CeCl}_3\text{-NaI}$ or $\text{CeCl}_3\text{-SnCl}_2$ systems produce the aldols.¹²

Protocol 1.

Reduction of 2-cyclopentenone by $\text{NaBH}_4\text{-CeCl}_3$ ^{6,7} (Scheme 12.5).

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 12.5

Equipment

- An Erlenmeyer flask (300 mL) equipped with a stirring bar.

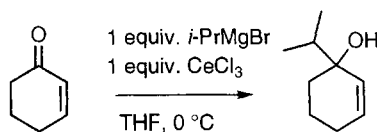
Materials

- Cerium chloride heptahydrate (FW 372.6), 19.0 g, 50 mmol irritant, hygroscopic
- 2-Cyclopentenone (FW 82.10), 4.1 g, 50 mmol
- Methanol, 120 mL flammable, toxic
- Sodium borohydride (FW 37.83), 1.9 g, 50 mmol flammable, corrosive

1. Add cerium chloride heptahydrate (50 mmol), 2-cyclopentenone (50 mmol), methanol (120 mL) and stirring bar to the flask. Stir the mixture until the cerium chloride is completely dissolved.
2. Immerse the flask in an ice bath, add sodium borohydride (50 mmol) in portions over 5 min. Hydrogen is evolved vigorously with elevation of temperature to 15–20°C.
3. Stir the mixture for an additional 10 min, and then concentrate the resulting white suspension to about 30 mL under reduced pressure.
4. Add water (100 mL) to dissolve inorganic salts and extract the mixture with ether (3×50 mL). Dry the combined extracts over MgSO_4 and concentrate to a volume of c. 10 mL. Distil the residue under ordinary pressure to collect 2-cyclopenten-1-ol (3.7 g, 88%) with 96% purity.

Protocol 2.**Preparation of organocerium reagents and reaction with carbonyl compounds: reaction of *i*-propylcerium reagent with 2-cyclohexene-1-one⁹ (Scheme 12.6)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 12.6****Equipment**

- A two-necked, round-bottomed flask (500 mL) fitted with a septum, a three-way stopcock, and a magnetic stirring bar. The three-way stopcock is connected to a vacuum/argon source.
- Vacuum/argon source.
- Syringes
- Mortar
- Oil bath

Materials

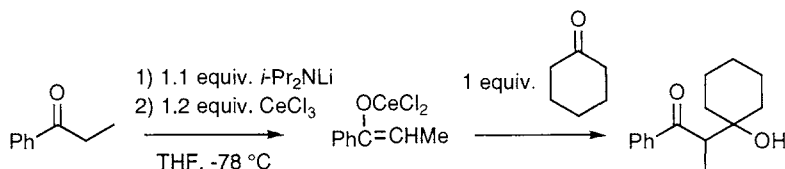
- Cerium chloride heptahydrate (FW 372.6), 18.0 g, 48 mmol
- *i*-Propylmagnesium bromide (FW 147.3), 1.8 M solution in THF, 27 mL, 48 mmol
- 2-Cyclohexene-1-one (FW 96.13), 3.85 g, 40 mmol
- Dry THF

irritant, hygroscopic**flammable, moisture sensitive****highly toxic****flammable corrosive**

1. Quickly finely grind cerium chloride heptahydrate to a powder in a mortar and place it in a two-necked round-bottomed flask. Immerse the flask in an oil bath and heat gradually to 135–140°C with evacuation (<0.5 mmHg). After 1 h at this temperature, completely dry the cerium chloride *in vacuo* by stirring at the same temperature for an additional hour. While the flask is still hot, introduce argon gas and cool the flask in an ice bath. Add dry THF (200 mL) all at once with vigorous stirring. Remove the ice bath and stir the suspension overnight under argon at room temperature.
2. Add a THF (27 mL) solution of *i*-PrMgBr (48 mmol) at 0°C, and stir the mixture at that temperature for 1.5 h.
3. Add a THF (10 mL) solution of 2-cyclohexene-1-one (40 mmol) dropwise over 10 min.
4. After 30 min at 0°C, add an aqueous solution of acetic acid (3%, 150 mL) and extract with ether (3 × 50 mL). Wash the combined extracts with brine, a solution of NaHCO₃, and brine. Dry the extract over MgSO₄ and concentrate under reduced pressure. Distil the desired product (4.5 g, 80%) as a fraction boiling at 83–85°C at 15 mmHg. This sample may contain 3% of 1,4-addition product as determined by capillary GLC analysis (Silicon OV-17).

Protocol 3.**Aldol reaction with cerium enolate: synthesis of 2-(1-hydroxycyclohexyl)-1-phenyl-1-propanone^{10,11} (Scheme 12.7)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 12.7****Equipment**

- A two-necked, round-bottomed flask (200 mL) fitted with a septum, a three-way stopcock, and a magnetic stirring bar. The three-way stopcock is connected to a vacuum/argon source
- Vacuum/argon source
- Syringes

Materials

- Di-isopropylamine (FW 101.2), 1.5 mL **flammable, corrosive**
- $n\text{-BuLi}$ (FW 64.06), 1.53 M solution in hexane, 7.2 mL, 11 mmol **pyrophoric, moisture-sensitive**
- Propiophenone (FW 134.2), 1.34 g, 10 mmol
- Cerium chloride, anhydrous (FW 246.5) suspension in THF (see Protocol 2) **irritant, hygroscopic**
- Cyclohexanone (FW 98.15), 0.982 g, 11 mmol **corrosive, toxic**
- Dry THF **flammable, corrosive**

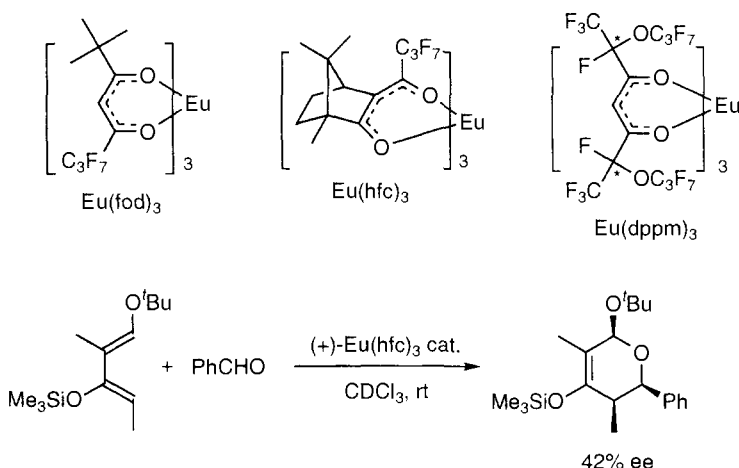
1. Flame dry the reaction vessel under dry argon. Quickly add a stirring bar while the flask is hot, and then refill the flask with argon through several vacuum cycles. Maintain a slightly positive argon pressure throughout this reaction.
2. Add THF (15 mL) and di-isopropylamine (1.5 mL) and cool the mixture to -78°C with a dry ice–acetone bath. Add dropwise a hexane solution of $n\text{-BuLi}$ (11 mmol) with stirring. After stirring for 10 min, add a THF (5 mL) solution of propiophenone (10 mmol).
3. After stirring for 1 h, add a THF (50 mL) suspension of anhydrous cerium chloride (12 mmol) (see Protocol 2) via a syringe and stir the mixture for 30 min at that temperature.
4. Add a THF (2 mL) solution of cyclohexanone (10 mmol).
5. Stirring for 30 min, then pour the mixture into an aqueous solution of acetic acid (3%, 50 mL). Extract with ether (3×50 mL), and then wash the combined extracts with brine, a solution of NaHCO_3 , and brine. Dry the extract

Protocol 3. Continued

over MgSO_4 and concentrate under reduced pressure. Purify the crude product by flash chromatography (silica gel, hexane–ethylacetate, 5:1) to give 2-(1-hydroxycyclohexyl)-1-phenyl-1-propane (1.67 g, 72%), $[\alpha]^{20}_{\text{D}} -5.3^\circ$ (c 1.2, EtOH).

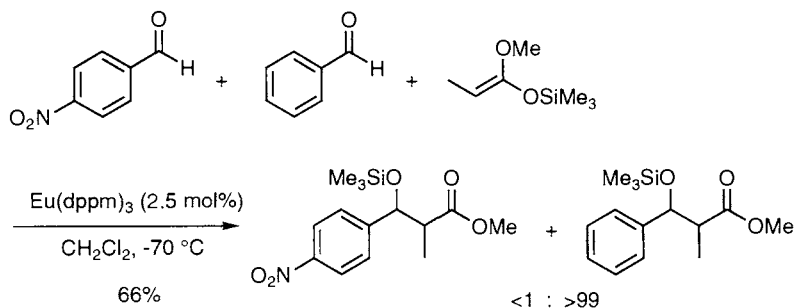
3. Europium(III) reagents

Although Eu(III) complexes such as $\text{Eu}(\text{fod})_3$ are familiar to organic chemists as NMR shift agents, the use of Eu complexes as synthetic reagents is a rather recent event. In 1983 Danishefsky *et al.* demonstrated the usefulness of (+)- $\text{Eu}(\text{hfc})_3$ as an asymmetric catalyst for the hetero Diels–Alder reaction (Scheme 12.8). Since then, a variety of Eu(III) complexes, including EuCl_3 and $\text{Eu}(\text{dppm})_3$, have been used as unique catalysts for different synthetic transformations.

**Scheme 12.8**

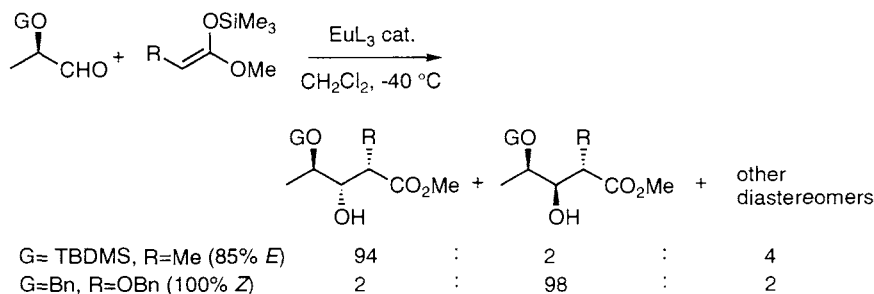
To date, europium(III) complexes have found many applications as efficient Lewis acid catalysts for Diels–Alder reactions,^{14,15} hetero Diels–Alder reactions,^{14,16–18} aldol reactions with ketene silyl acetals (KSA),^{19–25} Michael reactions of α,β -enones with KSA,²⁶ and cyanosilylation of aldehydes with silyl cyanides.²⁷ Compared with conventional Lewis acids such as TiCl_4 , SnCl_4 , and $\text{BF}_3 \cdot \text{OEt}_2$ (stoichiometric use), the europium catalyses are often quite unique in chemo- and stereoselectivities. For instance, $\text{Eu}(\text{dppm})_3$ is effective as the catalyst for aldol reactions between aldehydes and KSA, but not effective for the reactions between aldehydes and ketone-derived enol silyl ethers.²¹

Scheme 12.9 illustrates the unique ability of $\text{Eu}(\text{dppm})_3$ to discriminate the aldehydes, where *p*-nitrobenzaldehyde is less reactive than benzaldehyde. This reversal of reactivity is ascribed to preferential activation of benzaldehyde over *p*-nitrobenzaldehyde by the Eu catalyst. A similar molecular discrimination by $\text{Eu}(\text{fod})_3$ is achieved between ketones and α,β -enones in the reactions with KSA in which the Eu catalyst preferentially activates the latter in preference to the former.²²

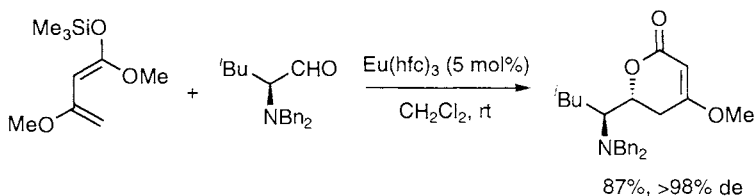


Scheme 12.9

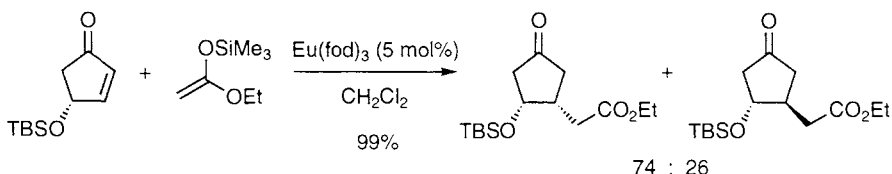
Of more significance is the unique stereocontrolling ability of the Eu-catalysts. For example, the $\text{Eu}(\text{fod})_3$ -catalysed aldol reactions of chiral α -alkoxy aldehydes with KSA provides high levels of diastereofacial selection and simple diastereoselections, where the stereochemistry of the major product is effectively modulated by changing the protecting group (G) through the molecular recognition (chelation vs. non-chelation) by the Eu catalyst (Scheme 12.10).^{23–25} A similar stereomodulation by the Eu catalyst is attained in the hetero Diels–Alder reactions of α -amino and α -alkoxy aldehydes (Scheme 12.11).^{28,29} Interestingly, the Michael addition of the 4-siloxy cyclopentenone with the acetate-derived KSA is shown to proceed with the sterically less favourable *syn* diastereofacial selectivity (Scheme 12.12).³⁰



Scheme 12.10



Scheme 12.11

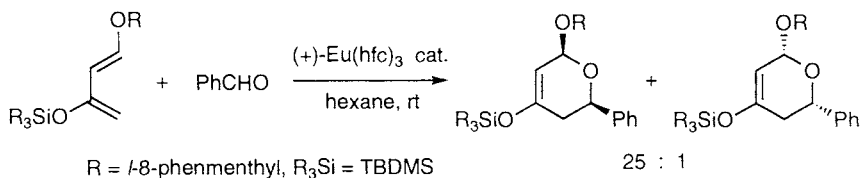


Scheme 12.12

As far as asymmetric catalysis by chiral europium complexes is concerned, modest enantioselectivities (up to 42% ee) are observed in the (+)- $\text{Eu}(\text{hfc})_3$ catalysed hetero Diels–Alder reactions (Scheme 12.8),¹³ whereas almost no asymmetric inductions (up to 10% ee) are observed in the (–)- $\text{Eu}(\text{dppm})_3$ -catalysed aldol reactions.²¹ More recently, the chiral europium complex, prepared *in situ* from $\text{Eu}(\text{OTf})_3$ and the disodium salt of a chiral diamine, has been reported to provide *c.* 40% ee in the aldol reactions of benzaldehyde with KSA.³¹ Thus, the development of asymmetric catalysis by chiral Eu complexes remains a challenge.

Protocol 4. Diels–Alder reaction¹⁷ (Scheme 12.13)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 12.13

Equipment

- A two-necked, round-bottomed flask fitted with a septum, a three-way stopcock, and a magnetic stirring bar. A three-way stopcock is connected to a vacuum/argon source
- Vacuum/argon source
- Syringes

Materials

- Benzaldehyde (FW 106.12), 234 mg, 2.20 mmol
- *l*-8-Phenmethoxy diene (FW 414.71), 912 mg, 2.20 mmol
- (+)-Eu(hfc)₃ (FW 1193.73), 131 mg, 0.11 mmol
- Dry Hexane, 12 mL
- Triethylamine (FW 101.19), 8 mL
- Methanol 4 mL

highly toxic, cancer suspect agent

hygroscopic
flammable, irritant
flammable, corrosive
flammable, toxic

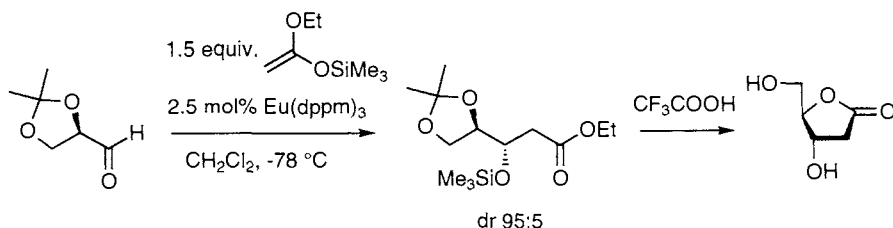
1. Add a solution of benzaldehyde (2.20 mmol) and *l*-8-phenmethoxy diene (2.20 mmol) in hexane (12 mL) to the flask. Cool the mixture to -20°C and add (+)-Eu(hfc)₃ (0.11 mmol). Keep the reaction mixture at between -10 and -20°C for 60 h.^a
2. Add triethylamine (8 mL) and methanol (4 mL) and allow the mixture to warm to room temperature. Remove the volatiles *in vacuo*, and pass the crude material through a plug of silica gel. Wash the plug with ethyl acetate. Concentrate the organics under reduced pressure to give enol silyl ether (1.10 g, 95%, dr 25:1). Crystallize this material from ethanol and recrystallize this residue from the mother liquors to give optically pure compound as white crystals (0.69 g, 60%); m.p. $59.5\text{--}60.8^{\circ}\text{C}$, $[\alpha]^{23}_{\text{D}} +47.3^{\circ}$ (c 1.1, CHCl_3).

^a After which time NMR analysis showed that no starting aldehyde or diene remained.

Protocol 5.

Eu(fod)₃-catalysed aldol reaction: preparation of optically pure 2-deoxy-D-ribonolactone²³ (Scheme 12.14)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 12.14

Equipment

- A two-necked, round-bottomed flask (30 mL) fitted with a septum, a three-way stopcock, and a magnetic stirring bar. The three-way stopcock is connected to a vacuum/argon source
- A round-bottomed flask (30 mL) fitted with a stirring bar
- Vacuum/argon source.
- Syringes

Protocol 5. Continued**Materials**

- D-Glyceraldehyde acetonide (FW 130.1), 130 mg, 1.0 mmol
- 1-Ethoxy-1-trimethylsiloxyethene (FW 160.3), 240 mg, 1.5 mmol
- (–)-Eu(dppm)₃ (FW 2072), 30 w/v% solution in CF₂ClCCl₂,^a 172 μL, 0.025 mmol
- Dry CH₂Cl₂

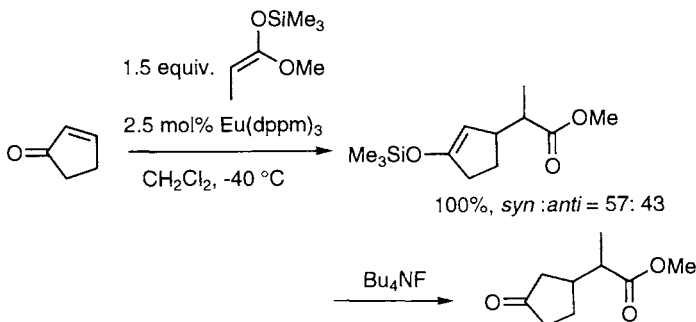
toxic, irritant

1. Flame dry the reaction vessel under dry argon. Quickly add a stirring bar while the flask is hot, and then refill the flask with argon through several vacuum cycles. Maintain a slightly positive argon pressure throughout this reaction.
2. Add a solution of D-glyceraldehyde acetonide (1 mmol) and 1-ethoxy-1-trimethylsiloxyethene (1.5 mmol) in CH₂Cl₂ (4 mL) to the flask with the aid of a syringe through the septum and then cool it to –78°C. Add a solution of (–)-Eu(dppm)₃ in CF₂ClCCl₂ (172 μL, 0.025 mmol) and then stir the mixture for 2 h at the same temperature.
3. Add saturated aqueous NH₄Cl solution and extract with ethyl acetate (3 × 20 mL). Wash the combined extracts with brine, dry over MgSO₄, and concentrate the crude product under reduced pressure (diastereomer ratio; *anti/syn* = 95:5).
4. Dissolve a crude product in THF (2 mL) and to this add trifluoroacetic acid (0.5 mL) and one drop of water.
5. After stirring for 1 day, add aqueous NaHCO₃ solution and extract with ethyl acetate (5 × 30 mL). Wash the combined extracts with brine and concentrate under reduced pressure. Purify the crude product by flash chromatography (silica gel; hexane–ethylacetate, 5:1) to give optically pure 2-deoxy-D-ribonolactone (0.118 g, 90%), [α]_D²⁰ 25.3° (c 1.2, EtOH).

^a Purchased from Dai-ichi Kagaku Yakuin.

Protocol 6.**(+)-Eu(dppm)₃-catalysed Michael reaction: reaction of a ketenesilylacetal with 2-cyclopentenone.²⁶ (Scheme 12.15)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 12.15****Equipment**

- A two-necked, round-bottomed flask (30 mL) fitted with a septum, a three-way stopcock, and a magnetic stirring bar. The three-way stopcock is connected to a vacuum/argon source
- Vacuum/argon source
- Syringes

Materials

- 2-Cyclopentenone (FW 82.10), 41 mg, 0.5 mmol
- (E)-1-[(trimethylsilyl)oxy]-1-methoxy-1-propene
- (+)-Eu(dppm)₃, 30 w/v% solution in CF₂ClCFCl₂^a
- Dry CH₂Cl₂
- Tetrabutylammonium fluoride (FW 261.5), 1.0 M solution in THF
- Dry THF

toxic, irritant
flammable, irritant
flammable, corrosive

1. Flame dry the reaction vessel under dry argon. Quickly add a stirring bar while the flask is hot, and then refill the flask with argon through several vacuum cycles. Maintain a slightly positive argon pressure throughout this reaction.
2. Add a solution of 2-cyclopentenone (0.5 mmol) in CH₂Cl₂ (1.5 mL) and a 30 w/v% solution of (+)-Eu(dppm)₃ (0.0125 mmol) in CF₂ClCFCl₂ to the flask with the aid of a syringe through the septum and then cool it to -40 °C. Add (E)-1-[(trimethylsilyl)oxy]-1-methoxy-1-propene (0.75 mmol) to the stirred solution.
3. After stirring for 1 h, pour the mixture into a aqueous saturated solution of NaHCO₃. Extract with ethyl acetate (3 × 30 mL). Wash the combined extracts with brine, dry over MgSO₄, and concentrate the crude product under reduced pressure.

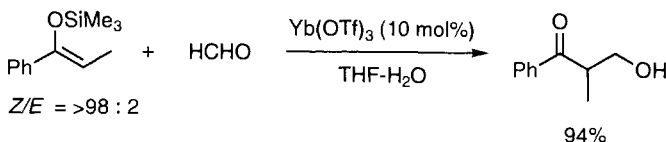
Protocol 6. Continued

4. Dissolve a crude product in THF (1 mL) and to this add a 1.0 M THF solution of Bu₄NF (1 mL) at room temperature. After stirring for several hours and monitor the disappearance of the silyl enol ether by TLC.
5. Pour the mixture into a water (20 mL) and extract with ethyl acetate (5 × 30 mL). Wash the combined extracts with brine and concentrate under reduced pressure. The crude product was purified by flash chromatography (silica gel, hexane–ethyl acetate) to give 3-[(1'-carbomethoxy)ethyl]-1-cyclopentanone (78 mg, 100%) as a diastereomer mixture (*syn:anti* = 57:43).

^a Purchased from Dai-ichi Kagaku Yakuhin.

4. Ytterbium(III) reagents

The most important of these reagents is ytterbium triflate, Yb(OTf)₃, which has found wide application as a catalyst for several C–C bond-forming reactions. The most notable is the Yb(OTf)₃-catalysed aldol reaction of aldehydes (or acetals) with enol silyl ethers which proceeds smoothly in dichloromethane and, more significantly, even in aqueous media (Scheme 12.16).^{32–36} It should be noted that a similar catalytic activity is also observed with Nd(OTf)₃, Gd(OTf)₃, and Lu(OTf)₃.

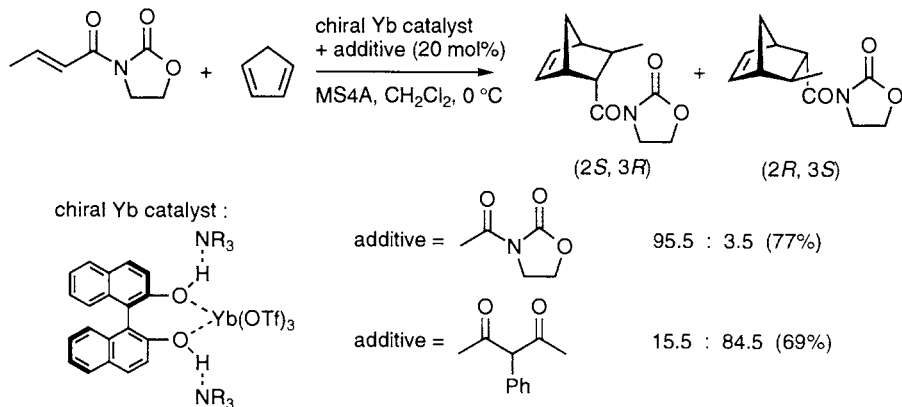


Scheme 12.16

In addition, Yb(OTf)₃ is also effective as a catalyst for the Michael reactions of α,β-enones with enol silyl ethers,³⁷ the Friedel–Crafts acylations with acid anhydrides,^{38,39} and the Diels–Alder reactions, wherein Yb(OTf)₃ is, however, often less active than Sc(OTf)₃. A major merit of the Yb-catalysed reactions is that the catalyst is stable in water and hence can be easily recovered from the aqueous layer and reused.³⁵

More significantly, the chiral Yb triflate, prepared *in situ* from Yb(OTf)₃, (*R*)-binaphthol, and a tertiary amine, is shown to serve as an efficient asymmetric catalyst for the Diels–Alder reaction (Scheme 12.17).⁴⁰ The most striking feature of the asymmetric catalysis is that the chiral Yb catalyst, when combined with 3-phenylacetylacetone as an additive (another achiral ligand), shows the opposite sense of enantioselection to that observed in the cases where the substrate itself or 3-acetyl-1,3-oxazolin-2-one acts as another ligand

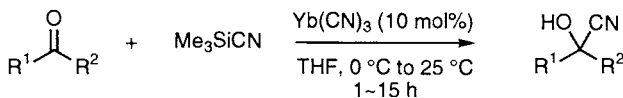
12: Lanthanide(III) Reagents



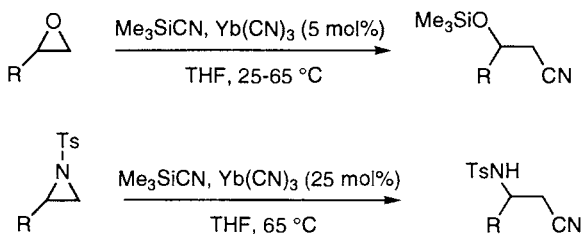
Scheme 12.17

as depicted in Scheme 12.17. The remarkable enantiocontrol by a proper choice of achiral ligands is explained in terms of the significant difference in complexation mode between the chiral catalytic species involved and the imide dienophile. A similar phenomenon is observed when other lanthanide triflates such as $\text{Lu}(\text{OTf})_3$, $\text{Tm}(\text{OTf})_3$, and $\text{Ho}(\text{OTf})_3$ are used in place of $\text{Yb}(\text{OTf})_3$.

Further notable ytterbium reagents are $\text{Yb}(\text{fod})_3$ as an efficient catalyst for the ene-type reaction of aldehydes with vinylic ethers,⁴¹ and $\text{Yb}(\text{CN})_3$ as an active catalyst for the cyanosilylation of carbonyls, including enolizable ketones and α,β -enones (Scheme 12.18),⁴² and for the regioselective ring opening of oxiranes and aziridines with Me_3SiCN (Scheme 12.19).^{43,44}

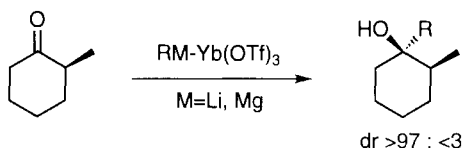


Scheme 12.18



Scheme 12.19

Ytterbium complexes have also found stoichiometric uses in organic synthesis. The alkylytterbium complex, generated *in situ* from $\text{Yb}(\text{OTf})_3$ or YbCl_3 and an alkyl-lithium or Grignard reagent, provides a unique and high stereoselectivities in the additions to ketones (Scheme 12.20).^{45,46}

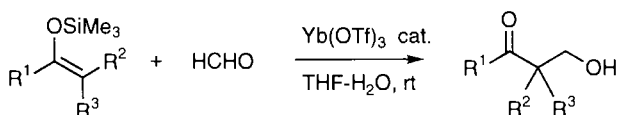


Scheme 12.20

Protocol 7.

$\text{Yb}(\text{OTf})_3$ -catalysed aldol reaction: reaction of an enol trimethylsilyl ether with formaldehyde. General procedure³⁶ (Scheme 12.21)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 12.21

Equipment

- A round-bottomed flask fitted with a stirring bar

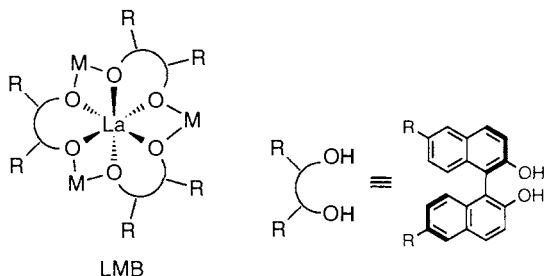
Materials

- Formaldehyde (FW 30.03), 40 wt. % solution in water, 1 mL **highly toxic, cancer-suspect agent**
- Enol silyl ether, 0.4 mmol
- $\text{Yb}(\text{OTf})_3$ (FW 620.3), 25 mg, 0.04 mmol **irritant, hygroscopic**
- THF **flammable, corrosive**

1. Add commercial formaldehyde solution (40% water solution, 1 mL) and THF (3 mL) to the flask.
2. Add $\text{Yb}(\text{OTf})_3$ (0.04 mmol, 10 mol%) and a THF (1 mL) solution of enol silyl ether (0.4 mmol) at room temperature, and then stir the mixture for 24 h at this temperature.
3. Remove the THF under reduced pressure. Add water and extract the product with CH_2Cl_2 . After the usual work-up, chromatograph the crude product on silica gel to yield the aldol product. $\text{Yb}(\text{OTf})_3$ was almost quantitatively recovered from the aqueous layer after removing water and could be reused.

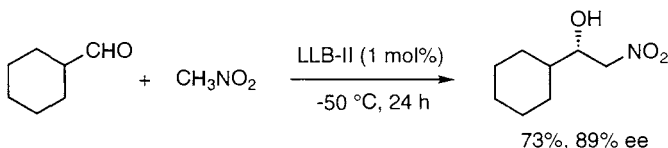
5. Lanthanum(III) reagents

Recently $\text{La}_3(\text{O}^t\text{Bu})_9$ has been found to serve as an efficient Lewis base catalyst for C–C bond-forming reactions such as the aldol reaction of α -chloro ketones with aldehydes, the nitroaldol (Henry) reactions of aldehydes with nitroalkanes, and the intramolecular aldol reactions of 1,5-diketones.⁴⁷ Based on these findings, Shibasaki's group has developed several (*R*)- or (*S*)-binaphthol (BINOL)-based chiral bimetallic lanthanum complexes, $\text{LaM}_3\text{tris}(\text{binaphthoxide})$ LMB ($\text{M} = \text{Li}, \text{Na}, \text{K}$) (Scheme 12.22), which serve as excellent asymmetric catalysts for the nitroaldol reaction, the hydrophosphonylation of aldehydes and imines, the Diels–Alder reaction, and the Michael reaction.⁴⁸



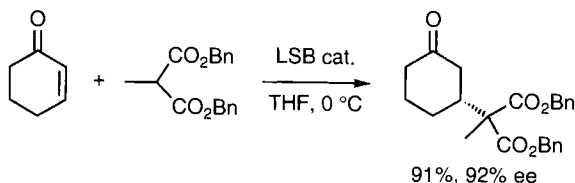
Scheme 12.22

Of special value is the enantioselective nitroaldol reaction catalysed by (*R*)-LLB, prepared *in situ* from either $\text{LaCl}_3 \cdot 7\text{H}_2\text{O}$, the dilithium salt of (*R*)-BINOL, and NaO^tBu or from $\text{La}(\text{O}^i\text{Pr})_3$ and (*R*)-BINOL, and *n*-BuLi, and then H_2O , which affords the nitroaldol in a high % ee (Scheme 12.23).^{49,50} Interestingly, the use of either the second generation LLB-II catalyst (LLB + LiOH) or the 6,6'-disubstituted BINOL-based LLB provides a higher % de and % ee in the reactions leading to a diastereomeric mixture of the nitroaldols.



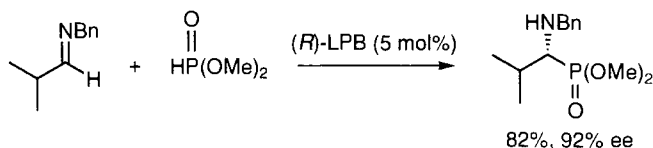
Scheme 12.23

LLB also serves as an efficient asymmetric catalyst for the Diels–Alder reaction⁴⁹ and the hydrophosphonylation of aldehydes,^{51–53} but not for the Michael reaction and the hydrophosphonylation of imines. Significantly, however, LSB and LPB are shown to work as an excellent asymmetric catalyst for the Michael reaction (Scheme 12.24)^{54,55} and the hydrophosphonylation of



Scheme 12.24

imines (Scheme 12.25),⁵⁶ respectively. It should be noted that similar chiral bimetallic complexes derived from other lanthanides such as EuMB, GdMB, and PrMB (M = Li, Na, K) often show a comparably high catalytic activity, but provides a much lower % ee in general.

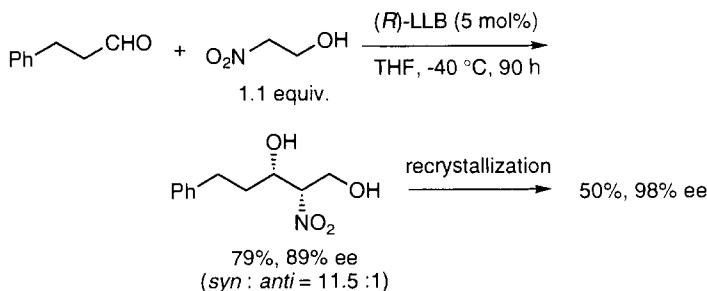


Scheme 12.25

Protocol 8.

Catalytic asymmetric nitroaldol reaction⁵⁰

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 12.26

Equipment

- A two-necked, round-bottomed flask fitted with a septum, a three-way stopcock, and a magnetic stirring bar. A three-way stopcock is connected to a vacuum/argon source
- Vacuum/argon source
- Syringes

Materials

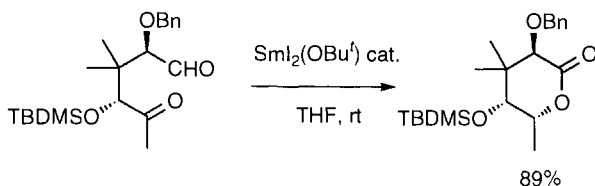
- (*R*)-(+)-1,1'-Bi-2-naphthol (FW 286.33), 6.5 g, 22.7 mmol irritant
- *n*-Butyllithium (FW 64.06), 1.60 M solution in hexanes, 28.4 mL, 45.4 mmol flammable, moisture sensitive
- $\text{LaCl}_3 \cdot 7\text{H}_2\text{O}$ (FW 371.38), 3.12 g, 8.40 mmol irritant, hygroscopic
- $\text{NaO}-t\text{-Bu}$ (FW 96.11), 242 mg, 2.52 mmol flammable, corrosive
- 2-Nitroethanol (FW 91.07), 19.1 mg, 0.21 mmol irritant
- Hydrocinnamaldehyde (FW 134.18), 25.5 mg, 0.19 mmol irritant
- Dry THF flammable, corrosive

1. Add (*R*)-(+)-1,1'-Bi-2-naphthol (22.7 mmol) to the flask. Dry at 50°C for 2 h under reduced pressure. Add dry THF (119 mL) under Ar,
2. Add a hexanes solution of *n*-BuLi (1.60 N, 28.4 mL, 45.4 mmol) at 0°C.
3. Charge a suspension of $\text{LaCl}_3 \cdot 7\text{H}_2\text{O}$ (8.40 mmol) in THF (100 mL) to the different flask. Sonicate for 30 min at room temperature. To this suspension, add the above-prepared solution of dilithium binaphthoxide and a THF solution of $\text{NaO}-t\text{-Bu}$ (0.52 N, 4.85 mL, 2.52 mmol), dropwise at room temperature.
4. Vigorously stir the mixture overnight at room temperature, and then for 48 h at 50°C. Stand the mixture at room temperature. The resulting supernatant may contain an optically active LLB complex (0.03 M).
5. Charge a THF solution of optically active LLB complex (0.03 M, 209 μL , 0.00627 mmol) to the flask. Add THF (500 μL) and cool the mixture to -40°C and stir for 30 min.
6. Add a THF (200 μL) solution of 2-nitroethanol (15 μL , 0.21 mmol), and stir the mixture for 30 min at the same temperature.
7. Add hydrocinnamaldehyde (25 μL , 0.19 mmol), and then stir the mixture for 90 h at -40°C).
8. Add a 1N HCl and then extract with AcOEt (30 mL); wash the organic layer with brine. Dry the extract over Na_2SO_4 and concentrate under reduced pressure. Purify the crude product by chromatography (Lobar Lichroeprep RI-8 prepac column, $\text{CH}_3\text{CN}-\text{H}_2\text{O}$ 1:1), to give (2*S*,3*S*)-2-nitro-5-phenyl-1,3-pentanediol (70%, *syn:anti* = 11:1, 91% ee). Determine the enantiomeric excess HPLC analysis using DAISEL CHIRALPAK AD (hexane: *i*-PrOH 9:1).

6. Other lanthanide(III) reagents

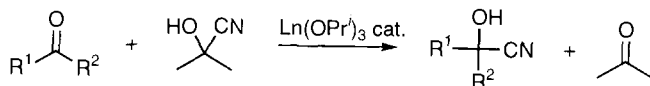
So far, the synthetic utilities of the four important classes of trivalent lanthanide reagents have been described. Outlined below are selected examples of the synthetic uses of other trivalent lanthanide complexes as catalysts.

A series of lanthanide isopropoxides, including $\text{La}(\text{OPr}^i)_3$, $\text{Ce}(\text{OPr}^i)_3$, $\text{Sm}(\text{OPr}^i)_3$, $\text{Gd}(\text{OPr}^i)_3$, have been found to exhibit an efficient catalytic activity in the Meerwein-Ponndorff-Verley (MPV) reductions. These catalysts



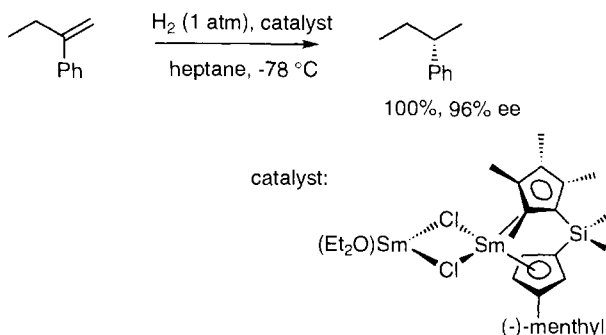
Scheme 12.27

are far more active than the existing catalysts such as $\text{Al}(\text{OPr}^i)_3$. Interestingly, $\text{Sm}(\text{OBu}^t)_2$, prepared *in situ* from SmI_2 and di-*t*-butylperoxide, works well as a catalyst for MPV reduction^{57–60} and also for the intramolecular Tishchenko reaction (Scheme 12.27).⁶¹ These lanthanide alkoxides also serve as an efficient catalyst for the transhydrocyanation of acetone cyanohydrin to aldehydes or ketones (Scheme 12.28).⁶²



Scheme 12.28

$\text{Lu}(\text{III})$ and $\text{Sm}(\text{III})$ hydride complexes having pentamethylcyclopentadienyl (Cp') ligands work as an extremely active homogeneous catalyst for the hydrogenation of terminal olefins in particular.^{63,64} Of special value is the asymmetric hydrogenation using a chiral samarium catalyst with a chiral auxiliary on the Cp ring (Scheme 12.29).⁶⁵



Scheme 12.29

Finally, some lanthanide complexes such as $(\text{Cp}'_2 \text{LnR})_2$ ($\text{Ln} = \text{La}, \text{Sm}, \text{Yb}, \text{Nd}, \text{Lu}$; $\text{R} = \text{H}, \text{Me}$) have been shown to be extremely active homogeneous catalysts for the polymerization of ethylene (not propylene), methyl metha-

crylate (MMA), ϵ -caprolactam, and δ -valerolactone.^{66,67} Of interest is the lanthanide-catalyzed polymerization of MMA which proceeds in a living polymerization fashion to produce polymers of high molecular weight ($>5 \times 10^5$) with a narrow polydispersity ($M_w/M_n = 1.02\text{--}1.03$) and a high syndiotacticity ($>95\%$).^{68,69}

References

1. Imamoto, T. *Lanthanides in Organic Synthesis*. Academic Press; London, **1994**.
2. Kagan, H. B.; Namy, J. L. *Tetrahedron* **1986**, *42*, 6573–6614.
3. Imamoto, T. *J. Synth. Org. Chem.* **1988**, *46*, 540–552.
4. Inanaga, J. *J. Synth. Org. Chem.* **1989**, *47*, 200–211.
5. Ho, T.-L. In *Organic synthesis by oxidation with metal compounds* (ed. M. J. Mijs, C. R. H. I. de Jonge). Plenum, New York, **1986**.
6. Luche, J.-L. *J. Am. Chem. Soc.* **1978**, *100*, 2226–2227.
7. Gemal, A. L.; Luche, J.-L. *J. Am. Chem. Soc.* **1981**, *103*, 5454–5459.
8. Bestmann, H. J.; Roth, D. *Synlett* **1990**, 751–753.
9. Imamoto, T.; Takiyama, N.; Makamura, K.; Hatajima, T.; Kamiya, Y. *J. Am. Chem. Soc.* **1989**, *111*, 4392–4398.
10. Imamoto, T.; Kusumoto, T.; Yokoyama, M. *Tetrahedron Lett.* **1983**, *24*, 5233–5236.
11. Nagasawa, K.; Kanbara, H.; Matsushita, K.; Ito, K. *Tetrahedron Lett.* **1985**, *26*, 6477–6480.
12. Fukuzawa, S.; Tsuruta, T.; Fujinami, T.; Sakai, S. *J. Chem. Soc. Perkin Trans. 1* **1987**, 1473–1477.
13. Bednarski, M.; Maring, C.; Danishefsky, S. *Tetrahedron Lett.* **1983**, *24*, 3451–3454.
14. Morrill, T. C.; Clark, R. A.; Bilobran, D.; Youngs, D. S. *Tetrahedron Lett.* **1975**, 397–400.
15. Ishar, M. P. S.; Wali, A.; Gandhi, R. P. *J. Chem. Soc. Perkin Trans. 1* **1990**, 2185–2192.
16. Bednarski, M.; Danishefsky, S. *J. Am. Chem. Soc.* **1983**, *105*, 3716–3717.
17. Bednarski, M.; Danishefsky, S. *J. Am. Chem. Soc.* **1986**, *108*, 7060–7067.
18. Quimpère, M.; Jankowski, K. *J. Chem. Soc. Chem. Commun.* **1987**, 676–677.
19. Takai, K.; Heathcock, C. H. *J. Org. Chem.* **1985**, *50*, 3247–3251.
20. Vougioukas, A. E.; Kagan, H. B. *Tetrahedron Lett.* **1987**, *28*, 5513–5516.
21. Mikami, K.; Terada, M.; Nakai, T. *J. Org. Chem.* **1991**, *56*, 5456–5459.
22. Hanyuda, K.; Hirai, K.; Nakai, T. *Synlett* **1997**, 31–32.
23. Mikami, K.; Terada, M.; Nakai, T. *Tetrahedron Asym.* **1991**, *2*, 993–996.
24. Terada, M.; Gu, J.-H.; Deka, D. C. Mikami, K.; Nakai, T. *Chem. Lett.* **1992**, 29–32.
25. Mikami, K.; Terada, M.; Nakai, T. *J. Chem. Soc. Chem. Commun.* **1993**, 343–345.
26. Terada, M.; Nakai, T.; Mikami, K. *Inorg. Chim. Acta* **1994**, *222*, 377–380.
27. Gu, J.-H.; Okamoto, M.; Terada, M.; Mikami, K.; Nakai, T. *Chem. Lett.* **1992**, 1169–1172.
28. Midland, M. M.; Afonso, M. M. *J. Am. Chem. Soc.* **1989**, *111*, 4368–4371.
29. Midland, M. M.; Koops, R. W. *J. Org. Chem.* **1990**, *55*, 4647–4650, 5058–5065.
30. Otera, J.; Fujita, Y.; Fukuzumi, S.; Hirai, K.; Gu, J.-H.; Nakai, T. *Tetrahedron Lett.* **1995**, *36*, 95–98.

31. Uotsu, K.; Sasai, H.; Shibasaki, M. *Tetrahedron Asymm.* **1995**, *6*, 71–74.
32. Kobayashi, S. *Chem. Lett.* **1991**, 2187–2190.
33. Kobayashi, S.; Hachiya, I. *Tetrahedron Lett.* **1992**, *33*, 1625–1628.
34. Kobayashi, S.; Hachiya, I.; Takahori, T. *Synthesis* **1993**, 371–373.
35. Kobayashi, S. *Synlett* **1994**, 689–701.
36. Kobayashi, S.; Hachiya, I. *J. Org. Chem.* **1994**, *59*, 3590–3596.
37. Kobayashi, S.; Hachiya, I.; Takahori, T.; Araki, M.; Ishitani, H. *Tetrahedron Lett.* **1992**, *33*, 6815–6818.
38. Kawada, A.; Mitsumura, S.; Kobayashi, S. *J. Chem. Soc., Chem. Commun.* **1993**, 1157–1158.
39. Mikami, K.; Kotera, O.; Motoyama, Y.; Sakaguchi, H.; Maruta, M. *Synlett* **1996**, 171–172.
40. Kobayashi, S.; Ishitani, H. *J. Am. Chem. Soc.* **1994**, *116*, 4083–4084.
41. Deaton, M. V.; Ciufolini, M. A. *Tetrahedron Lett.* **1993**, *34*, 2409–2412.
42. Matsubara, S.; Takai, T.; Utimoto, K. *Chem. Lett.* **1991**, 1447–1450.
43. Matsubara, S.; Onishi, H.; Utimoto, K. *Tetrahedron Lett.* **1990**, *31*, 6209–6212.
44. Matsubara, S.; Kodama, T.; Utimoto, K. *Tetrahedron Lett.* **1990**, *31*, 6379–6380.
45. Molander, G. A.; Burkhardt, E. R.; Weinig, P. *J. Org. Chem.* **1990**, *55*, 4990–4991.
46. Utimoto, K.; Nakamura, A.; Matsubara, S. *J. Am. Chem. Soc.* **1990**, *112*, 8189–8190.
47. Sasai, H.; Suzuki, T.; Arai, S.; Arai, T.; Shibasaki, M. *J. Am. Chem. Soc.* **1992**, *114*, 4418–4420.
48. Shibasaki, M.; Sasai, H.; Arai, T. *Angew. Chem. Int. Ed. Engl.* **1997**, *36*, 1236–1256.
49. Arai, T.; Yamada, Y. M. A.; Yamamoto, N.; Sasai, H.; Shibasaki, M. *Chem. Eur. J.* **1996**, *2*, 1368–1372.
50. Sasai, H.; Watanabe, S.; Shibasaki, M. *Enantiomer* **1988**, *2*, 267–271.
51. Yokomatsu, T.; Yamagishi, T.; Shibuya, S. *Tetrahedron Asymm.* **1993**, *4*, 1783–1784.
52. Rath, N. P.; Spilling, C. D. *Tetrahedron Lett.* **1994**, *35*, 227–230.
53. Sasai, H.; Bougauchi, M.; Arai, T.; Shibasaki, M. *Tetrahedron Lett.* **1997**, *38*, 2717–2720.
54. Sasai, H.; Arai, T.; Shibasaki, M. *J. Am. Chem. Soc.* **1994**, *116*, 1571–1572.
55. Sasai, H.; Arai, T.; Satow, Y.; Houk, K. N.; Shibasaki, M. *J. Am. Chem. Soc.* **1995**, *117*, 6194–6198.
56. Sasai, H.; Arai, S.; Tahara, Y.; Shibasaki, M. *J. Org. Chem.* **1995**, *60*, 6656–6657.
57. Namy, J. L.; Soupe, J.; Collin, J.; Kagan, H. B. *J. Org. Chem.* **1984**, *49*, 2045–2049.
58. Collin, J.; Namy, J. L.; Kagan, H. B. *Nouv. J. Chim.* **1986**, *10*, 229–232.
59. Lebrun, A.; Namy, J. L.; Kagan, H. B. *Tetrahedron Lett.* **1991**, *32*, 2355–2358.
60. Okano, T.; Matsuoka, M.; Konishi, H.; Kiji, J. *Chem. Lett.* **1987**, 181–184.
61. Uenishi, J.; Masuda, S.; Wakabayashi, S. *Tetrahedron Lett.* **1991**, *32*, 5097–5100.
62. Ohno, H.; Mori, A.; Inoue, S. *Chem. Lett.* **1993**, 375–378.
63. Mauermann, H.; Swepston, P. N.; Marks, T. J. *Organometallics* **1985**, *4*, 200–202.
64. Jeske, G.; Lauke, H.; Mauermann, H.; Schumann, H.; Marks, T. J. *J. Am. Chem. Soc.* **1985**, *107*, 8111–8118.
65. Conticello, V. P.; Brard, L.; Giardello, M. A.; Tsuji, Y.; Sabat, M.; Stern, C. L.; Marks, T. J. *J. Am. Chem. Soc.* **1992**, *114*, 2761–2762.

12: Lanthanide(III) Reagents

66. Jeske, G.; Lauke, H.; Mauermann, H.; Swepston, P. N.; Schumann, H.; Marks, T. J. *J. Am. Chem. Soc.* **1985**, *107*, 8091–8103.
67. Watson, P. L.; Parshall, G. W. *Acc. Chem. Res.* **1985**, *18*, 51–56.
68. Yasuda, H.; Yamamoto, H.; Yokota, K.; Miyake, S.; Nakamura, A. *J. Am. Chem. Soc.* **1992**, *114*, 4908–4910.
69. Yasuda, H.; Furo, M.; Yamamoto, H.; Nakamura, A.; Miyake, S.; Kibino, N. *Macromolecules* **1992**, *25*, 5115–5116.

Other transition metal reagents: chiral transition-metal Lewis acid catalysis for asymmetric organic synthesis

HISAO NISHIYAMA and YUKIHIRO MOTOYAMA

1. Introduction

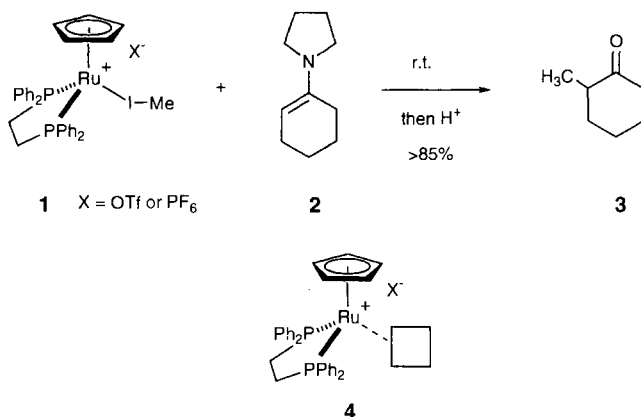
Transition-metal Lewis acid catalysts can supply one or occasionally two vacant sites as Lewis acids for substrates possibly to meet the 18-electron rule for transition metals, and moreover they need not change their valency during the reactions. The unique vacant sites prepared by releasing one or two ligands from the characteristic configurations ML_n ($n = 4-6$, M = metal, L = ligand) consists of different kinds of steric circumstances. Therefore, design of their ligands with consideration not only of stereochemistry but also of valency of metals has been a remarkably attractive subject to develop new Lewis acid catalysts especially for asymmetric synthesis.

The complexes of the early transition metals such as the family of Sc, Ti, V, Cr, or Mn (where d^n , $n \leq 6$), being electron-poor, can easily supply vacant sites as Lewis acids as mentioned in the early chapters. In contrast, those of the late transition metals such as Fe, Co, Ni, or Cu etc., being electron-rich, should be converted into their cationic species to act as acids, such as $[ML_{n-1}]^+$ ($n = 4-6$), accompanied by certain conjugate bases as counter anions.

2. Activation of haloalkanes and carbonyl compounds

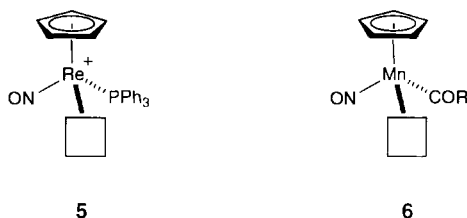
Simple haloalkanes have been found to bind to transition-metal complexes via the halogen lone pair. The co-ordinating iodomethane can act as a methylating reagent which methylated several nucleophiles faster than free iodomethane. For example, Crabtree *et al.* reported that the reaction of the iodomethane complex of Cp-Ru **1** exhibited methylation of the cyclohexanone

enamine **2** to selectively give a 2-methylcyclohexanone **3** in high yield by C-alkylation (Scheme 13.1),^{1,2} although free enamines accompany *N*-alkylation products. Upon co-ordination to the cationic ruthenium(II), the methyl carbon atom of MeI increases electrophilicity attributed to activation by the cationic ruthenium Lewis acid as **4**.



Scheme 13.1

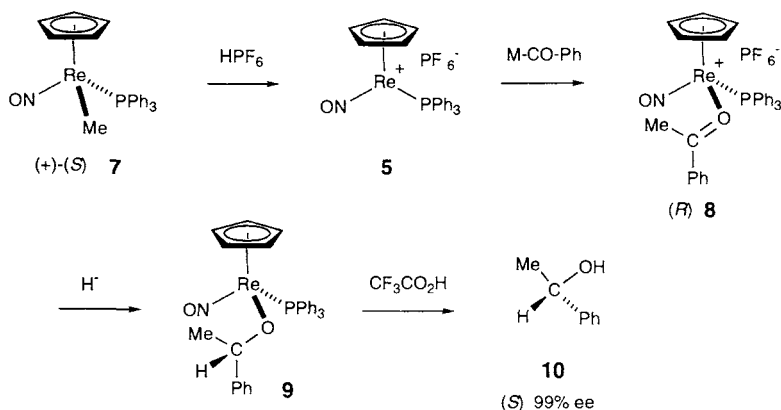
Similar haloalkane complexes of Cp-rhenium have also been reported.^{3,4} The cationic complex [CpRe(NO)(PPh₃)(IME)]⁺BF₄⁻ can methylate triphenylphosphine at room temperature to produce a phosphonium salt in *c.* 90% yield.⁵ The CpRe(NO)(PPh₃)(vacant)⁺ species **5** (Scheme 13.2) can act as a Lewis acid and activate the carbon-halogen bonds. Moreover, the Cp-Re species **5** maintain the chirality-at-metal during several reactions.⁵ In the sense of chirality-at-metal, a similar observation was disclosed by Brunner in 1974 that the manganese complex CpMn(NO)(COR)(PAR₃) exhibits a dissociative substitution by PAR'₃ with retention of the configuration via the intermediate **6**.⁶⁻⁸



Scheme 13.2

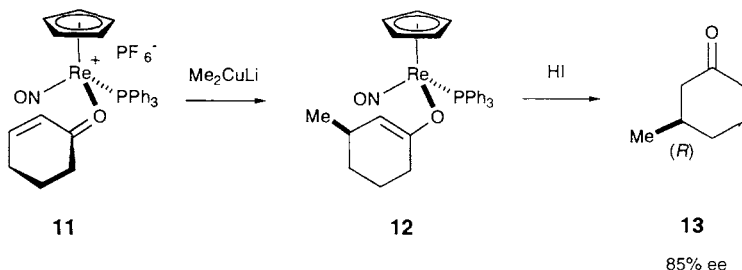
The σ -complex of acetophenone-rhenium was prepared from the methyl complex **7** by treatment with HPF₆ followed by addition of acetophenone via

the unsaturated species **5** (Scheme 13.3).⁹ The acetophenone complex **8** when analysed by X-ray showed that the rhenium atom preferentially binds *syn* to the smaller methyl substituent. The methyl is also directed towards the smallest NO ligand, where the *re*-face of the ketone is opened. Therefore, hydride reduction of the complex (+)-(*R*)-**8** was converted into the corresponding alkoxide complex **9** followed by acidification to produce the optically active 1-phenylethanol **10** in 99% ee (*S*) (Scheme 13.3).



Scheme 13.3

Similarly, the α,β -unsaturated ketone moiety of the chiral rhenium–enone complex (+)-(*R*)-**11** was attacked by dimethyl cuprate followed by acidification to give β -methyl cyclohexanone **13** in 85% ee via the adduct **12** (Scheme 13.4).¹⁰



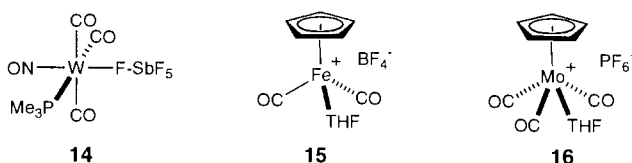
Scheme 13.4

Thus, stoichiometric modes of activation of haloalkanes and carbonyl compounds have been realized in both non-asymmetric and asymmetric reactions by use of the transition-metal complexes as Lewis acids. Many other examples are referred to in the literatures.^{5–10}

3. Catalytic asymmetric Diels–Alder reactions

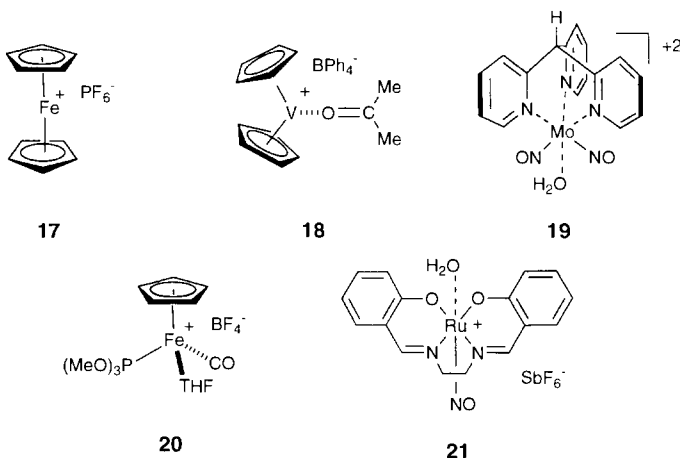
3.1 Achiral transition-metal catalysts

Low oxidation state transition-metal Lewis acids were first reported as catalysts of the Diels–Alder reactions by Hersh *et al.*^{11,12} They originally found the tungsten nitrosyl complex $[\text{W}(\text{PMe}_3)(\text{CO})_3(\text{NO})]^+(\mu\text{-F-SbF}_5)^-$ **14** (Scheme 13.5) as a catalyst of butadiene or cyclopentadiene and α,β -unsaturated enones at 0°C . They suggested that the mode of catalysis is likely due to activation of the α,β -unsaturated enone by simple $\eta^1(\sigma)$ -carbonyl co-ordination on the basis of X-ray analysis of the tungsten–acrolein adduct. They also compared the related cationic fragments of the corresponding η^1 -carbonyl complexes, such as $\text{Cp}(\text{CO})_2\text{Fe}^+$ **15** and $\text{Cp}(\text{CO})_3\text{Mo}^+$ **16**, to well-known transition-metal Lewis acids.¹² The catalytic activity of these reagents for the Diels–Alder reactions was compared; **14** > **15** > **16**. Furthermore, the order of Lewis acidity was measured on the basis of ^1H NMR chemical shifts of crotonaldehyde on co-ordination; Lewis acids (relative power), BBr_3 (1.00) > AlCl_3 (0.82) > BF_3 (0.77) > TiCl_4 (0.66) > **14** (0.62) > **16** (0.47) > AlEt_3 (0.44) > **15** (0.36).



Scheme 13.5

Kelly *et al.*¹³ reported ferrocenium ion, **17** (Scheme 13.6) acting as a catalyst for the Diels–Alder reactions of cyclopentadiene and α,β -unsaturated enones



Scheme 13.6

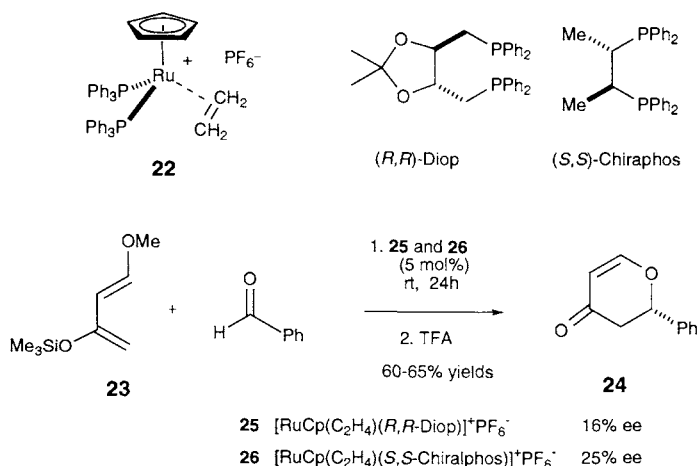
in good yields at 0–20°C. The possibility that metallocenes may serve as Lewis acids can be inferred from the vanadocenium complex of acetone $(\text{Cp})_2\text{V}(\eta^1\text{-O}=\text{CMe}_2)^+$, **18**. However, other metallocenes, such as zirconocene dichloride, titanocene dichloride, cobaltocenium ion, and ferrocene, did not exhibit the catalytic activity of the Diels–Alder reaction as Lewis acids.

Other transition-metal Lewis acids for the Diels–Alder reactions have been reported: the doubly charged complex $[\text{HC}(\text{py})_3\text{Mo}(\text{NO})_2(\text{solvent})]^{+2}(\text{SbF}_6)_2$, **19**,¹⁴ $(\text{Cp})(\text{CO})[\text{P}(\text{OMe})_3]\text{Fe}^+$, **20**,¹⁵ and $[\text{Ru}(\text{salen})(\text{NO})(\text{H}_2\text{O})]\text{SbF}_6$, **21**.^{16,17}

3.2 Chiral transition-metal catalysts

Many of transition-metal complexes, as described above, can sufficiently activate dienophiles by co-ordination of the carbonyl groups. The reactions proceed smoothly under mild conditions to provide high diastereoselectivities and enantioselectivities. Therefore, much effort has been devoted to creating a new asymmetric Diels–Alder reaction with the combination of transition-metals and chiral auxiliaries. Many original chiral auxiliaries have been developed for this purpose.^{18,19}

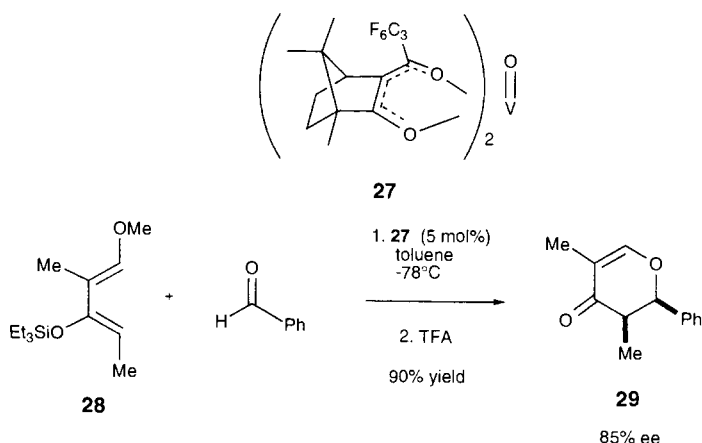
Faller *et al.*²⁰ found that the Cp–Ru cationic complex $[\text{CpRu}(\text{PPh}_3)_2(\text{C}_2\text{H}_4)]\text{PF}_6$, **22** (5 mol%) catalysed the hetero-Diels–Alder reaction of Danishefsky's diene **23** and benzaldehyde at room temperature to give the adduct **24** in 78% yield (Scheme 13.7). The catalytically active species generated by dissociation of ethylene from **22**, activate benzaldehyde reacting easily with the diene **23**. The optically active bidentate phosphines, Diop and Chiralphos, also entered to this reaction. However, the enantioselectivities were not high as expected, 16% ee with the Ru–Diop catalyst, **25**, and 25% ee with the Ru–Chiralphos catalyst, **26**. More recently, Ghosh *et al.*²¹ improved



Scheme 13.7

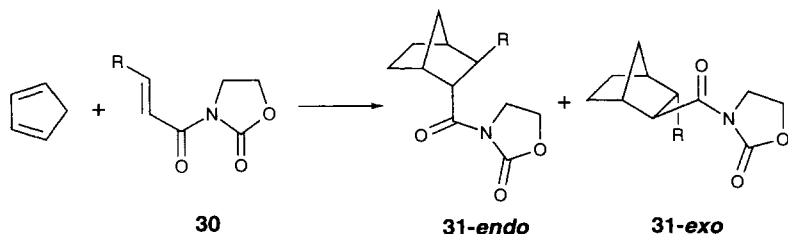
the hetero-Diels–Alder reaction of diene **23** with glyoxylates in the presence of chiral copper–bisoxazoline catalysts to attain 70% ee.

Optically active oxovanadium(IV) complexes with camphor-derived diketono ligands have been investigated for the asymmetric hetero-Diels–Alder reaction.²² The oxovanadium complex, **27**, (5 mol%) effectively catalysed the reaction of the diene **28** and benzaldehyde at -78°C for *c.* 1 day to give the adduct **29** in 90% yield and in 85% ee (Scheme 13.8).



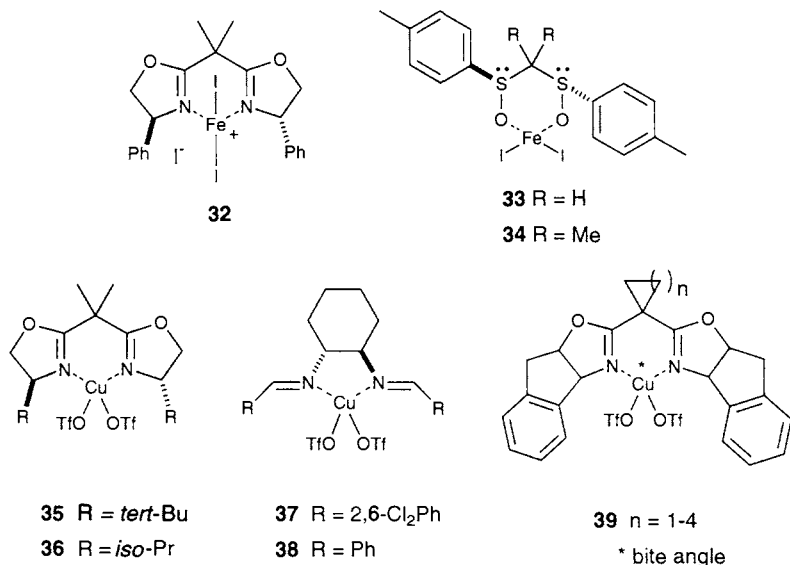
Scheme 13.8

As a bidentate dienophile, the acylimide **30** was introduced to the Diels–Alder reaction by Narasaka in 1986 (Scheme 13.9).²³ To capture this component, the reaction may preferentially need an acid catalyst providing two vacant sites. The octahedral structure of transition-metal complexes can supply this demand, as shown in the chiral iron–bisoxazoline complex, **32**, developed by Corey *et al.*²⁴ (Protocol 1) and the iron–bisulfonate complexes, **33** and **34** (Scheme 13.10) reported by Khier *et al.*²⁵ (Table 13.1). The iron–bisoxazoline, **32** (10 mol%) catalysed the reaction of the acylimide **30** and



Scheme 13.9

13: Other transition metal reagents



Scheme 13.10

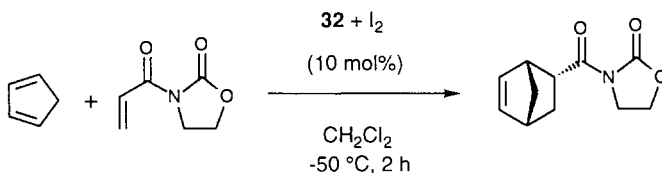
Table 13.1. Asymmetric Diels–Alder reaction with the acylimide **30** and cyclopentadiene with the catalysts **32–39**

| Catalyst | mol% | R= | °C, h | yield(%) | ratio of <i>endo:exo</i> | % ee of <i>endo</i> |
|---------------|------|----|-----------|----------|-----------------------------|------------------------|
| 32 | 10 | H | −50, 15 | — | 99:1 | 86 |
| 33 | 10 | H | −50, 5 | 74 | 68:32 | 36 |
| 34 | 10 | H | −50, 5 | 78 | 78:22 | 56 |
| 35 | 10 | H | −78, 18 | 86 | 98:2 | 98 |
| 36 | 10 | H | −78–50, 4 | 93 | 96:4 | 58 |
| 35 | 10 | Me | −15, 30 | 85 | 96:4 | 97 |
| 37 | 9 | H | −78, 36 | 87 | 80:20 | 92 |
| 38 | 9 | H | −78, 36 | — | 80:20 | 85 |
| 37 | 9 | Me | −10, 30 | 90 | 65:35 | 83 |
| 38 n=1 | 10 | H | −70, — | 90 | 96:1 | 98 |
| 39 n=4 | 10 | H | −50, — | 90 | 26:1 | 83 |

cyclopentadiene at −50°C to give an *endo:exo* ratio of 99:1 and 86% ee for the *endo* isomer of **31**. It was postulated for both reactions that the iron catalysts **32–34** may supply two vacant sites of the equatorial–equatorial or the equatorial–axial.

Protocol 1.**Asymmetric Diels–Alder reaction of 3-acryloyl-1,3-oxazolin-2-one with cyclopentadiene catalysed by chiral bis(oxazoline)-Fe(III) complex²⁴ (Scheme 13.11)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

**Scheme 13.11****Equipment**

- Magnetic stirrer
- Syringes
- Oil bath
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)
- Cooling bath with dry ice–acetone

Materials

- | | |
|--|-------------------------------------|
| • Cyclopentadiene, 0.68 mL, 8.19 mmol | flammable, toxic |
| • 3-Acryloyl-1,3-oxazolin-2-one, 385 mg, 2.73 mmol | flammable, toxic |
| • 2,2-Bis[2-[4(S)-phenyl-1,3-oxazolinyl]]propane, 109 mg, 0.326 mmol | |
| • Powdered iron, 15.2 mg, 0.272 mmol | moisture-sensitive |
| • Iodine, 173.1 mg, 0.680 mmol | toxic, corrosive |
| • Acetonitrile, 4 mL | flammable, lachrymator |
| • Dichloromethane, 7 mL | harmful by inhalation |
| • Triethylamine, 0.1 mL | flammable, corrosive |
| • Ether | flammable, irritant |
| • Pentane | flammable, irritant |
| • Sodium sulfite | moisture-sensitive, irritant |
| • Copper acetate | irritant, hygroscopic |

1. Prepare a catalyst solution of powdered iron (15.2 mg, 0.272 mmol) and iodine (104 mg, 0.408 mmol) in acetonitrile (2 mL) under argon atmosphere. After stirring for 1 h at 40°C, add a solution of 2,2-bis[2-[4(S)-phenyl-1,3-oxazolinyl]]propane (109 mg, 0.326 mmol) in acetonitrile (2 mL) at 23°C and stir for 1 h at 40°C. Evaporated under reduced pressure to afford a dark viscous oil, and add dichloromethane (7 mL).
2. After cooling the catalyst solution to $-78^{\circ}C$, add iodine (69.1 mg, 0.282 mmol), 3-acryloyl-1,3-oxazolin-2-one (385 mg, 2.73 mmol) and cyclopentadiene (0.68 mL, 8.19 mmol).
3. After stirring for 2 h at $-50^{\circ}C$, add triethylamine (0.1 mL) and dilute with ether–pentane to the reaction mixture.

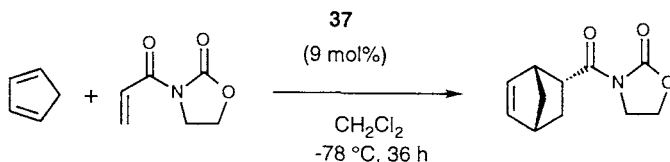
4. Wash the dichloromethane solution with aqueous sodium sulfite and copper acetate. Dry the organic layer and then remove the organic solvent using a rotary evaporator.
5. Apply the crude product to a flash silica gel column with 1:1 mixture of *n*-hexane and ethyl acetate to obtain pure product in 95% yield (539 mg, 96% *endo*, 82.2% ee).
6. Determine the enantiomeric excess by HPLC analysis (Daicel CHIRALCEL OD, 10% *i*-PrOH in hexane, flow rate 1 mL/min), $t_R = 21.4$ min (2*S*), 23.5 min (2*R*).

The most remarkable progress came from the use of the catalysts, **35** and **36** derived from copper(II) triflate and chiral bisoxazoline.²⁶ The reaction of the acylimide **30** and cyclopentadiene in the presence of the catalyst **35** (10 mol%) proceeds at -78°C for 5 h to give the adduct **31** in 86% with 98:2 ratio of the *endo:exo* isomers and 98% ee for the *endo* isomer. Similarly, the copper(II)-bisimine catalyst, **37** and **38**, exhibited high activity at -78°C to give 92% ee (Protocol 2). The sense of asymmetric induction with these copper catalysts can be interpreted by assuming the reaction via a square-planar Cu(II)(chiral ligand)(acylimide) complex, rather than a tetragonal one, which is attacked on the least hindered side by the diene. In a similar system to the copper catalysts, the copper catalysts of new bisoxazolines, **39**, were synthesized to show that the larger the bite angle of the bisoxazolines, the higher the enantioselectivity up to 98 from 82.²⁸

Protocol 2.

Asymmetric Diels–Alder reaction of 3-acryloyl-1,3-oxazolin-2-one with cyclopentadiene catalysed by chiral bis(imine)-Cu(II) complex²⁷ (Scheme 13.12)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 13.12

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice–acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Protocol 2. Continued**Materials**

- Cyclopentadiene, 0.68 mL, 8.19 mmol flammable, toxic
- 3-Acryloyl-1,3-oxazolin-2-one, 385 mg, 2.73 mmol flammable, toxic
- Bis(imine), 43 mg, 0.1 mmol
- Copper(II) triflate, 32.5 mg, 0.09 mmol corrosive, moisture-sensitive
- Dichloromethane, 2 mL harmful by inhalation
- Ether flammable, irritant

1. Prepare a catalyst solution of bis(imine) (43 mg, 0.1 mmol) and copper(II) triflate (32.5 mg, 0.09 mmol) in dichloromethane (2 mL) at room temperature under nitrogen atmosphere and stir for 5 h.
2. After cooling of the catalyst solution to -78°C , add 3-acryloyl-1,3-oxazolin-2-one (141 mg, 1 mmol) and cyclopentadiene (0.83, 10 mmol).
3. After stirring for 36 h at -78°C , dilute the reaction mixture with ether and filter through a small plug (1 cm \times 2 cm) of silica gel. Wash the plug three times with ether (totally 30 mL) : 87% yield, 80% *endo*, 92% ee.

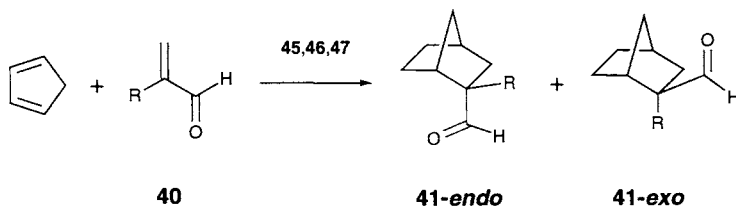
The non-asymmetric Diels–Alder reaction of α,β -unsaturated aldehydes have also been intensively examined with the transition-metal Lewis acids **14–21** as described above (Scheme 13.13). The first asymmetric reaction with chiral transition-metal Lewis acids was realized by Kündig *et al.* in 1994.²⁹ They demonstrated that the chiral iron–bisphosphite catalyst, **42**, (5 mol%), which itself has one acrolein molecule, exhibits catalytic activity for the reaction of acrolein derivatives and cyclopentadiene at -30°C giving in 84–99% ee (Table 13.2). The reaction of methacrolein with **42** as a catalysts is described in Protocol 3 (Scheme 13.14).²⁹

Table 13.2. Asymmetric Diels–Alder reaction with acrolein derivatives and cyclopentadiene with the catalysts **42–44**

| Catalyst | R= | $^{\circ}\text{C}$, h | yield(%) | ratio of <i>endo:exo</i> | % ee of major |
|-------------------------|----|------------------------|------------------|--------------------------|---------------|
| 42 | H | -30 , 16 | 46 | 62:38 | 84 |
| 42 | Me | -20 , 20 | 62 | 3:97 | 90 |
| 42 | Et | -20 , 20 | 55 | 2:98 | 94 |
| 42 | Br | -40 , 16 | 87 | 5:95 | 95 |
| 43a ^a | H | -20 , 120 | >95 ^b | 96:4 | 85 |
| 43b ^a | H | -20 , 18 | >95 ^b | 94:6 | 85 |
| 43a ^a | Me | -20 , 120 | >95 ^b | 4:96 | 85 |
| 43b ^a | Me | -40 , 8 | >95 ^b | 3:97 | 92 |
| 43a ^a | Br | -40 , 60 | >95 ^b | 3:97 | 87 |
| 43b ^a | Br | -78 , 12 | >95 ^b | 2:98 | 96 |
| 44 | Me | -78 , 24 | 60 | 2:98 | 60 |
| 44 | Br | -78 , 24 | 93 | 2:98 | 68 |

^a **43** (5 mol%).

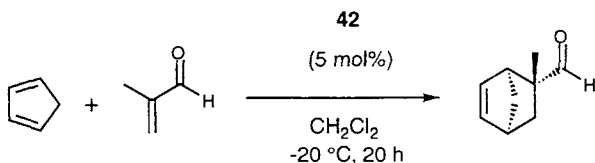
^b Conversion.



Scheme 13.13

Protocol 3.**Asymmetric Diels–Alder reaction of methacrolein with cyclopentadiene catalysed by chiral phosphite-Fe(II) complex²⁹ (Scheme 3.14)**

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 13.14

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice–acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Materials

- | | |
|--|------------------------------|
| • Phosphite-Fe-acrolein complex, 110 mg, 0.10 mmol | flammable, toxic |
| • Cyclopentadiene, 170 μL , 2.0 mmol | flammable, toxic |
| • Methacrolein, 170 μL , 2.0 mmol | flammable, toxic |
| • 2,6-Di- <i>tert</i> -butylpyridine, 12 μL , 0.05 mmol | irritant |
| • Dichloromethane, 2 mL | harmful by inhalation |
| • <i>n</i> -Hexane | flammable, irritant |
| • Ether | flammable, irritant |

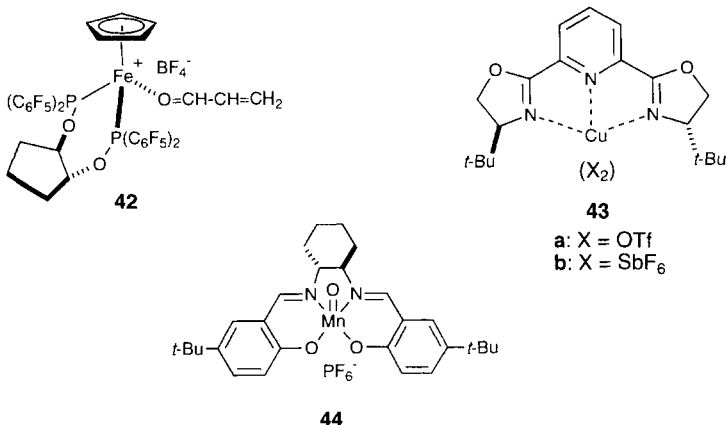
1. Prepare a catalyst solution of phosphite-Fe-acrolein complex (110 mg, 0.10 mmol) and 2,6-di-*tert*-butylpyridine (12 μL , 0.05 mmol) in dichloromethane (2 mL) at -40°C under argon atmosphere.
2. Add methacrolein (170 μL , 2.0 mmol) and stir for 15 min, then add cyclopentadiene (170 mL, 2.0 mmol) slowly.
3. After stirring for 20 h at -20°C , add hexane to precipitate the catalyst and wash three times with the same solvent.
4. Expose the combined organic phases to air for removing traces of residual

Protocol 3. Continued

- catalyst by oxidation. Filter through a pad of Celite and then remove the organic solvent with a rotary evaporator.
5. Apply the crude product to a flash silica gel column with *n*-hexane–ether (10:1) to obtain pure product in 62% yield (174 mg, 97% *exo*, 90% ee).
 6. Determine the *exo/endo* ratio by ^1H NMR; δ 9.40 (*endo*), 9.70 (*exo*).^a
 7. Determine the enantiomeric excess by capillary GC using PEG-25M, column temperature 80 °C; t_R = 36.4 (*exo-2R*), 38.1 min (*exo-2S*), 25.5, 33.1 min (*endo*), after converting the adduct to the corresponding (*R,R*)-2,4-dimethyl-1,3-dioxane. A mixture of the Diels–Alder adduct (10–20 mg), (*R,R*)-2,4-pentanediol (1.5 equiv), triethyl orthoformate (1.1 equiv), and *p*-toluene-sulfonic acid (1–2 mg) in dry benzene (1 mL) was stirred at room temperature for several hours (TLC check).^a

^aFuruta, K.; Shimizu, S.; Miwa, Y.; Yamamoto, H. *J. Org. Chem.* **1989**, *54*, 1481–1483. Determine the *exo/endo* ratio by capillary GC using PEG-20M, column temperature 100 °C; t_R = 8.7 min (*exo*), 10.3 min (*endo*).

Evans *et al.*³⁰ introduced the copper(II)-bis(oxazolinyl)pyridine (pybox) catalysts, **43**, (Scheme 13.15) for the same purpose, especially to find the importance of the counterion effect. They anticipated that the Cu–pybox has only one accessible co-ordination site for the carbonyl dienophiles, such as acrolein derivatives, because pybox is a tridentate ligand.^{31,32} The addition of methacrolein and cyclopentadiene with the Cu–pybox-*tb*, **43a**, (X = OTf) (5 mol%) at –20 °C for 120 h gave the adducts in the ratio of the *exo:endo*, 96:4 and 85% ee for the *exo* isomer (Table 13.2). However, the reaction with the Cu–pybox-*tb*, **43b** (X = SbF₆) proceeds only for 8 h to give 95% yield of



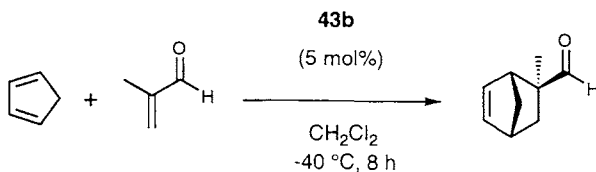
Scheme 13.15

the *exo*-isomer with 92% ee (Protocol 4) (Scheme 13.16). The same tendency was observed for bromoacrolein. The same counterion acceleration effect for the Diels–Alder reaction of the acylimide **30** with the Cu–bisoxazoline **35** ($X = \text{OTf} \Rightarrow \text{SbF}_6$) as also been shown.³⁰ The SbF_6 complex of **35** gave high yields of 96% for the less reactive β -substituted acylimides **30** ($R = \text{Ph}$ and Cl) with 95–96% ees.

Protocol 4.

Asymmetric Diels–Alder reaction of methacrolein with cyclopentadiene catalysed by chiral bis(oxazolinyl)pyridine–Cu(II) complex³⁰ (Scheme 13.16)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 13.16

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice–acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Materials

- | | |
|---|---|
| <ul style="list-style-type: none"> • Cyclopentadiene, 158 mg, 2.4 mmol • Methacrolein, 140 mg, 2.0 mmol • <i>tert</i>-Butyl[pyridine-bis(oxazoline)], 33 mg, 0.10 mmol • Copper bromide, 22 mg, 0.10 mmol • Silver hexafluoroantimonate, 69 mg, 0.20 mmol • Dichloromethane, 4 mL | <p>flammable, toxic</p> <p>flammable, corrosive</p> <p>irritant, hygroscopic</p> <p>corrosive, hygroscopic</p> <p>harmful by inhalation</p> |
|---|---|

1. Prepare a catalyst solution of powdered copper bromide (22 mg, 0.10 mmol), silver hexafluoroantimonate (69 mg, 0.20 mmol) and *tert*-butyl[pyridine-bis(oxazoline)] (33 mg, 0.10 mmol) in dichloromethane (4 mL). After stirring for 6 h, filter through a plug of cotton to give a clear blue–green solution.
2. After cooling of the catalyst solution to -78°C , add methacrolein (140 mg, 2.0 mmol) and cyclopentadiene (158 mg, 2.4 mmol): >95% yield, 97% *exo*, 92% ee.

Protocol 4. Continued

3. Determine the *exo/endo* ratio by capillary GC (DB-1701, 110°C, 5 psi), t_R = 5.40 (*exo*), 6.01 min (*endo*). Determine the enantiomeric excess by capillary GC (DB-1701, 110°C, 5 psi) after the adduct was converted into the corresponding (*R,R*)-2,4-dimethyl-1,3-dioxane, t_R = 29.89 (2*S*), 30.05 min (2*R*).

Very interestingly, optically active oxo(salen)manganese(V) complex **44** serves as a Lewis acid to catalyse the Diels–Alder reaction of bromoacrolein and cyclopentadiene in 93% yield (98:2 = *exo:endo*) with 68% ee for the *exo* isomer.³³

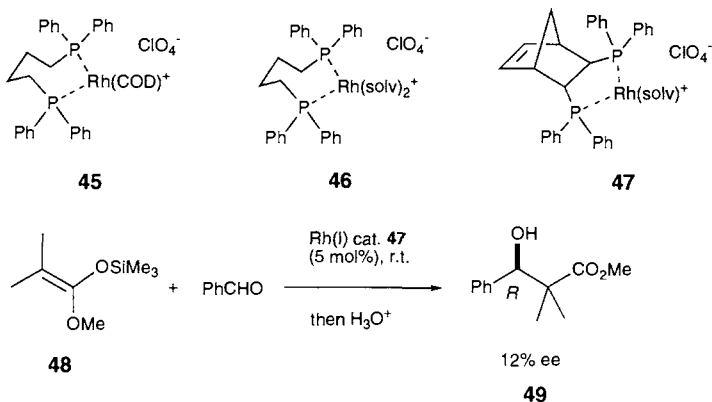
4. Asymmetric aldol condensations

A crossed aldol reaction of trimethylsilyl enol ethers and certain aldehydes catalysed by achiral cationic rhodium–diphosphine catalyst was first reported by Sato *et al.*³⁴ The reaction needs a relatively high temperature, 100°C for 15 h, but 2 mol% of $[(\text{COD})\text{Rh}(\text{DPPB})]^+\text{ClO}_4^-$ **45** (COD = 1,5-cyclooctadiene and DPPB = diphenylphosphinobutane) (Scheme 13.17) catalyses the reaction of $\text{Me}_3\text{SiOC}(\text{CH}_3)=\text{CH}_2$ and *n*-hexanal resulting in 74% yield of the corresponding adduct. They showed no catalytic activity of $\text{HRh}(\text{PPh}_3)_n$ ($n = 3$ and 4), $\text{HClRu}(\text{PPh}_3)_3$, and $\text{PdCl}_2(\text{CH}_3\text{CN})_2$ for that purpose, and they proposed a rhodium enolate as an intermediary.

Reetz *et al.*³⁵ reported similar results of the aldol reaction with more active cationic rhodium–diphosphine complex, $[(\text{solvent})_n\text{Rh}(\text{DPPB})]^+$, **46**, which was generated by treating the corresponding $[(\text{COD})\text{Rh}(\text{DPPB})]^+$ complex under a hydrogen atmosphere. The catalysed reaction of a ketene silylacetal, $\text{Me}_3\text{SiOC}(\text{OMe})=\text{CMe}_2$, **48**, and benzaldehyde with the rhodium catalysts (5 mol%) resulted in 81% yield of the adduct **49** for only 2 h at 22°C. It was pointed out that the mechanism of the aldol reaction may involve co-ordination of aldehyde molecules at the metal leading to activation of the aldehydes, but they did not deny the metal–enolate intermediate. They also applied the aldol reaction to an asymmetric version. The chiral Rh–Norphos cationic complex **47** (5 mol%) catalyses the aldol condensation of the ketene silylacetal **48** and benzaldehyde in good yields (>75%) but in a low enantioselectivity (Scheme 13.17).

Bosnich and colleagues has developed the achiral cationic ruthenium-based complex, $[\text{Ru}(\text{salen})(\text{NO})(\text{H}_2\text{O})]^+\text{SbF}_6^-$, **21** as a powerful transition-metal Lewis acid catalyst for the crossed-aldol reaction at 25°C in nitromethane solution, as well as the Diels–Alder reaction with the same catalyst described in the preceding section.^{17,36} The reaction of the ketene silylacetal **48** and benzaldehyde with **21** (1 mol%) for only 3 min resulted in over 90% reaction (Protocol 5) (Scheme 13.18).

13: Other transition metal reagents

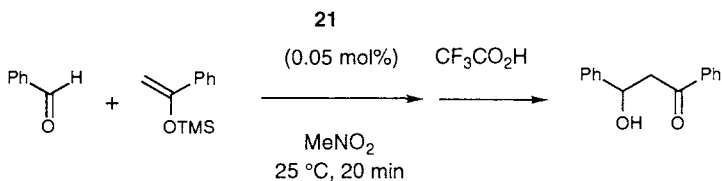


Scheme 13.17

Protocol 5.

Mukaiyama crossed-aldol reaction of 1-phenyl-1-[(trimethylsilyl)oxy]ethylene with benzaldehyde catalysed by [Ru(salen)(NO)H₂O]SbF₆ Complex³⁶ (Scheme 13.18)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 13.18

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice-acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Materials

- [Ru(salen)(NO)H₂O]SbF₆, 0.31 mg, 0.48 μmol
- Benzaldehyde, 104 mL, 1 mmol
- 1-Phenyl-1-[(trimethylsilyl)oxy]ethylene, 204 μL , 1 mmol
- Trifluoroacetic acid, 100 μL , 0.59 mmol
- Nitromethane, 700 μL , 2 mL
- Water, 3 mL
- Dichloromethane

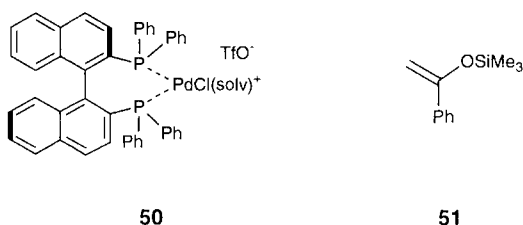
toxic, cancer-suspect agent
moisture-sensitive, irritant
corrosive, toxic
flammable

toxic, irritant

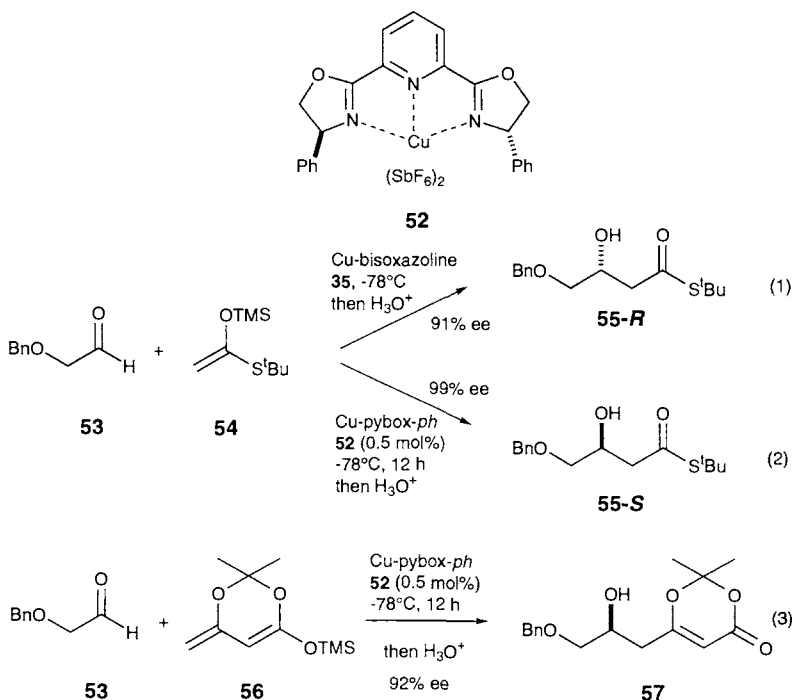
Protocol 5. Continued

1. Prepare a catalyst solution of $[\text{Ru}(\text{salen})(\text{NO})\text{H}_2\text{O}]\text{SbF}_6$,¹⁶ (0.31 mg, 0.48 μmol) in nitromethane (700 μL).
2. Add benzaldehyde (104 μL , 1 mmol) and 1-phenyl-1-[(trimethylsilyl)oxy]ethylene (204 μL , 1 mmol).
3. After stirring for 20 min at 25°C, dilute with nitromethane (2 mL) and stir with trifluoroacetic acid (100 μL) for 15 min.
4. Add water (3 mL) and stir vigorously for 15 min, then dilute with dichloromethane.
5. Extract the organic layer with dichloromethane and concentrate under reduced pressure.
6. Take up in dichloromethane and pass through a short column of Florisil, and remove the eluent solvent under reduced pressure to give the product, $\text{PhCOCH}_2\text{CH}(\text{OH})\text{Ph}$, (0.21 g, 92% yield) as an oil, pure by ^1H NMR.

The catalytic asymmetric aldol reaction with chiral palladium catalysts was reported by Shibasaki.^{37,38} The cationic complex $\text{PdCl}[(R)\text{-Binap}]^+$, **50** (Scheme 13.19) in DMF- H_2O catalysed the condensation of the acetophenone silyl enol ether **51** and benzaldehyde at 23°C to give 96% yield of the adduct with 71–73% ee. However, they concluded that the reaction does not involve a palladium Lewis acid catalyst, and that the reaction is the first example of the asymmetric aldol reaction via $\text{Pd}(\text{II})$ enolate species.

**Scheme 13.19**

A remarkable system was discovered using copper(II) complexes as Lewis acid catalysts, which activates α -alkoxy aldehydes through bidentate coordination.^{39,40} Evans *et al.* applied the cationic complexes of the copper(II)-bisoxazoline, **35** (OTf and SbF_6) and the copper(II)-pybox-*ph*, **52**, as catalysts (Scheme 13.20). The α -(benzyloxy)acetaldehyde, **53**, and the silylketene acetal, **54**, reacted in the presence of $\text{Cu}(\text{II})(\text{OTf})_2$ -bisoxazoline-*tb*-(*S,S*), **35** (5 mol%) at -78°C to give the (*R*)-adduct **55-R** in 91% ee (Scheme 13.20(1)). In contrast, the combination catalyst **53** of $\text{Cu}(\text{II})(\text{SbF}_6)_2$ (5 mol%) and pybox-*ph*-(*S,S*) gave the (*S*)-adduct **55-S** in 99% ee (Protocol 6) (Scheme



Scheme 13.20

13.20(2). Upon optimization, 0.5 mol% of the catalyst **53** catalysed the reaction for 12 h giving more than 94% yield and 92–98% ees for the several silyl enoethers, for example, from **56** to **57** in 92% ee (Scheme 13.20(3)).

Protocol 6.

Asymmetric mukaiyama-aldol reaction of (benzyloxy)acetaldehyde with 1-*tert*-butylthio-[(trimethylsilyl)oxy]ethene catalysed by chiral bis(oxazolinyl)pyridine-Cu(II) complex (Scheme 13.20^{39,40})

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.

Equipment

- Magnetic stirrer
- Syringes
- Cooling bath with dry ice–acetone
- Schlenk flask (10 mL)
- Vacuum/inert gas source (argon)

Materials

- Phenyl[pyridine-bis(oxazoline)], 9.2 mg, 0.025 mmol
- Copper chloride, 3.4 mg, 0.025 mmol
- Silver hexafluoroantimonate, 17.2 mg, 0.05 mmol

**irritant, hygroscopic
corrosive, hygroscopic**

Protocol 6. Continued

- Dichloromethane, 4 mL harmful by inhalation
- (Benzyloxy)acetaldehyde, 82.1 mg, 0.50 mmol
- 1-*tert*-Butylthio-[(trimethylsilyl)oxy]ethene, 122.7 mg, 0.60 mmol
- 1 M HCl toxic
- Tetrahydrofuran flammable, irritant

1. Prepare a catalyst solution of powdered copper chloride (3.4 mg, 0.025 mmol), silver hexafluoroantimonate (17.2 mg, 0.05 mmol) and phenyl[pyridine-bis(oxazoline)] (9.2 mg, 0.025 mmol) in dichloromethane (4 mL). After stirring for 4 h at room temperature, filter through a plug of cotton.
2. Add (benzyloxy)acetaldehyde (82.1 mg, 0.50 mmol) and 1-*tert*-butylthio-[(trimethylsilyl)oxy]ethene (122.7 mg, 0.60 mmol) at -78°C .
3. After stirring for 12 h at -78°C , filter the reaction mixture through silica, then hydrolyse silyl ether with 1 M HCl in tetrahydrofuran to give pure product in 100% yield (99% ee).
4. Determine the enantiomeric excess by HPLC analysis (Daicel CHIRALCEL OD-H).

Absolute configuration was assigned by comparison of optical rotation, see; Mikami, K.; Matsukawa, S. *J. Am. Chem. Soc.* **1994**, *116*, 2363–2364.

References

1. Kulawiec, R. J.; Crabtree, R. H. *Organometallics* **1988**, *7*, 1891.
2. Burk, M. J.; Segmuller, B.; Crabtree, R. H. *Organometallics* **1987**, *6*, 2241.
3. Winter, C. H.; Arif, A. M.; Gradysz, J. A. *J. Am. Chem. Soc.* **1987**, *109*, 7560.
4. Winter, C. H.; Veal, W. R.; Garner, C. M.; Arif, M. A.; Gradysz, J. A. *J. Am. Chem. Soc.* **1989**, *111*, 4766.
5. Fernández, J. M.; Gladysz, J. A. *Organometallics* **1989**, *8*, 207.
6. Brunner, H.; Aclasis, J.; Langer, M.; Steger, W. *Angew. Chem., Int. Ed. Engl.* **1974**, *13*, 810.
7. Brunner, H. *Adv. Organomet. Chem.* **1980**, *18*, 151.
8. Boone, B. J.; Klein, D. P.; Seyler, J. W.; Méndez, N. Q.; Arif, A. M.; Gladysz, J. A. *J. Am. Chem. Soc.* **1996**, *118*, 2411.
9. Fernández, J. M.; Emerson, K.; Larsen, R. D.; Gradysz, J. A. *J. Chem. Soc. Chem. Commun.* **1988**, 37.
10. Wang, Y.; Gladysz, J. A. *J. Org. Chem.* **1995**, *60*, 903.
11. Honeychuck, R. V.; Bonnesen, P. V.; Farahi, J.; Hersh, W. H. *J. Org. Chem.* **1987**, *52*, 5293.
12. Bonnesen, P. V.; Ruckett, C. L.; Honeychuck, R. V.; Hersh, W. H. *J. Am. Chem. Soc.* **1989**, *111*, 6070.
13. Kelly, T. R.; Maity, S. K.; Meghani, P.; Chandrakumar, N. S. *Tetrahedron Lett.* **1989**, *30*, 1357.
14. Faller, J. W.; Ma, Y. *J. Am. Chem. Soc.* **1991**, *113*, 1579.

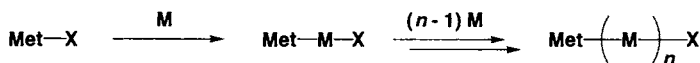
15. Olson, A. S.; Seitz, W. J.; Hossain, M. M. *Tetrahedron Lett.* **1991**, 32, 5299.
16. Odenkirk, W.; Reingold, A. L.; Bosnich, B. *J. Am. Chem. Soc.* **1992**, 114, 6392.
17. Hollis, T. K.; Odenkirk, W.; Robinson, N. P.; Whela, J.; Bosnich, B. *Tetrahedron* **1993**, 49, 5415.
18. Kagan, H. B.; Riant, O. *Chem. Rev.* **1992**, 92, 1007.
19. Maruoka, K.; Yamamoto, H. In *Catalytic asymmetric synthesis* (ed. I. Ojima). VCH, New York, 1993; Chapter 9, p 413.
20. Faller, J. W.; Smart, C. J. *Tetrahedron Lett.* **1989**, 30, 1189.
21. Ghosh, A. K.; Mathivanan, P.; Cappiello, J.; Krishnan, K. *Tetrahedron Asymm.* **1996**, 7, 2165.
22. Togni, A. *Organometallics* **1990**, 9, 3106.
23. Narasaka, K.; Inoue, M.; Okada, N. *Chem. Lett.* **1986**, 1109.
24. Corey, E. J.; Imai, N.; Zhang, H.-Y. *J. Am. Chem. Soc.* **1991**, 113, 728.
25. Khier, N.; Fernández, I.; Alcudta, F. *Tetrahedron Lett.* **1993**, 34, 123.
26. Evans, D. A.; Miller, S. J.; Lectka, T. *J. Am. Chem. Soc.* **1993**, 115, 6460.
27. Evans, D. A.; Lectka, T.; Miller, S. J. *Tetrahedron Lett.* **1993**, 34, 7027.
28. Davies, I. W.; Gerena, L.; Castonguay, L.; Senanayake, C. H.; Larsen, R. D.; Verhoeven, T. R.; Reider, P. J. *J. Chem. Soc., Chem. Commun.* **1996**, 1753.
29. Kündig, E. P.; Bourdin, B.; Bernardinelli, G. *Angew. Chem. Int. Ed. Engl.* **1994**, 33, 1856.
30. Evans, D. A.; Murry, J. A.; von Matt, P.; Norcross, R. D.; Miller, S. J. *Angew. Chem. Int. Ed. Engl.* **1995**, 34, 798.
31. Nishiyama, H.; Kondo, M.; Nakamura, T.; Itoh, K. *Organometallics* **1991**, 10, 500.
32. Nishiyama, H.; Itoh, Y.; Sugawara, Y.; Matsumoto, H.; Aoki, K.; Itoh, K. *Bull. Chem. Soc. Jpn.* **1995**, 68, 1247.
33. Yamashita, Y.; Katsuki, T. *Synlett.* **1995**, 829.
34. Sato, S.; Matsuda, I.; Izumi, Y. *Tetrahedron Lett.* **1986**, 27, 5517.
35. Reetz, M. T.; Voungioulas, A. E. *Tetrahedron Lett.* **1987**, 28, 793.
36. Odenkirk, W.; Whelan, J.; Bosnich, B. *Tetrahedron Lett.* **1992**, 33, 5729.
37. Sodeoka, M.; Ohrai, K.; Shibasaki, M. *J. Org. Chem.* **1995**, 60, 2648.
38. Kiyooka, S.; Tsutsumi, T.; Maeda, H.; Kaneko, Y.; Isobe, K. *Tetrahedron Lett.* **1995**, 36, 6531.
39. Evans, D. A.; Murry, J. A.; Kozlowski, M. C. *J. Am. Chem. Soc.* **1996**, 118, 5814.
40. Evans, D. A.; Kozlowski, M.; Tedrow, J. S. *Tetrahedron Lett.* **1996**, 37, 7481.

Lewis acid-assisted anionic polymerizations for synthesis of polymers with controlled molecular weights

TAKUZO AIDA and DAISUKE TAKEUCHI

1. Introduction

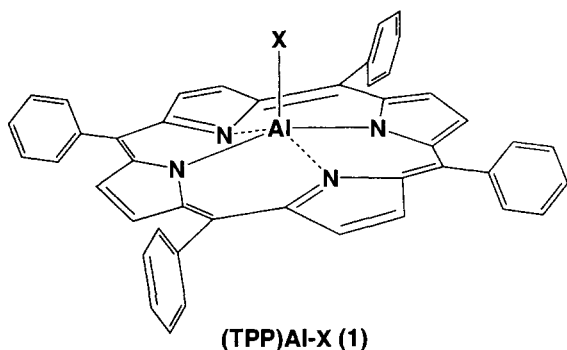
Unlike some naturally occurring macromolecules such as enzymes, synthetic polymers, except those obtained by stepwise approach, are of broad molecular weight distributions (MWD). In other words, they are mixtures of macromolecules with different molecular weights. Since 'molecular weight' is an essential factor affecting fundamental properties of polymer materials, it is important to develop a method for the synthesis of a polymer with a desired molecular weight with a narrow MWD. MWD is caused by the heterogeneity of the growth of polymer chains. Scheme 14.1 shows a schematic diagram of



Scheme 14.1

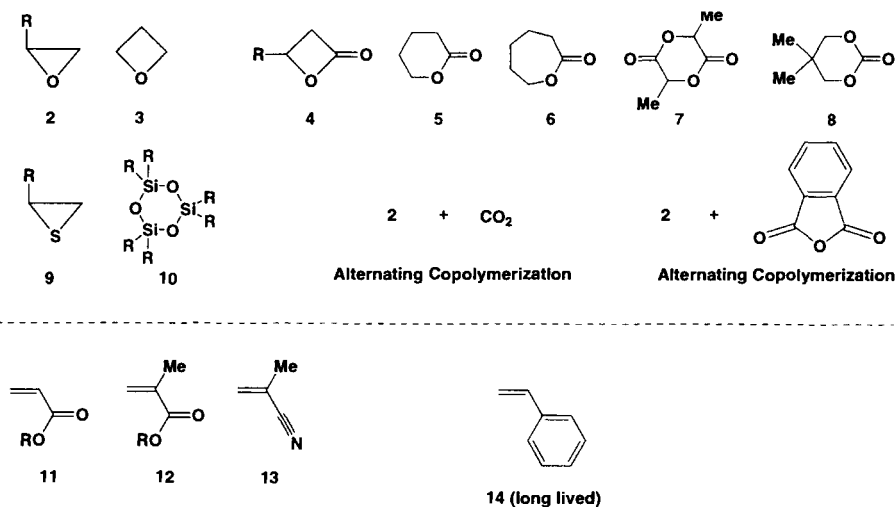
addition polymerization of an unsaturated monomer (M) with Met-X as initiator, where the chain growth of a polymer molecule starts by the reaction of Met-X with M to generate an active species (Met-M-X) (initiation step), followed by repeated additions of M to Met-M-X (propagation step), to furnish a higher molecular weight polymer (met-(M)_n-X). In this case, if (1) the initiation is much faster than the propagation, and (2) the propagation proceeds uniformly with respect to all growing polymer molecules (Met-(M)_n-X), a polymer with a uniform molecular weight should be formed. However, polymerization is generally accompanied by side reactions such as termination and chain transfer reactions, which lead to irreversible deactivation of the growing species. Since these side reactions interfere with the uniform growth of polymer chains, broadening of polymer MWD results.

In 1956, the first example of the formation of a narrow MWD polymer was discovered by Szwarc.¹ In his basic study on electron transfer reactions from organometallic compounds to unsaturated compounds, he noticed that mixing of styrene with sodium naphthalide resulted in the formation of a polymer, which is of very narrow MWD, as indicated by the ratio of weight- to number-average molecular weights (M_w/M_n) close to unity.² Upon addition of a fresh feed of styrene to this system after the complete consumption of the first feed, second-stage polymerization ensues, resulting in further growth of all growing polymer molecules. Polymerization with this character is named 'living polymerization', since the growth pattern of polymer can be viewed as analogous to the growth of a biological organism.



In order to achieve 'living polymerization', it is essential to develop well-behaved initiators, since the initiator affects the relative rate of initiation to propagation, and the potential for side reactions during chain growth. Aida and Inoue *et al.*³ have discovered that some metalloporphyrin complexes such as aluminium porphyrins ((TPP)Al-X, **1**; TPP = 5, 10, 15, 20-tetraphenylporphinato) serve as excellent initiators for living anionic polymerization, where the most characteristic feature is their exceptionally wide applicability for a variety of monomers (Scheme 14.2).⁴ With **1** as initiator, the polymerization proceeds by the insertion of a monomer (M) into the Al-X bond in **1** to give a (TPP)Al-M-X (initiation step), which reacts with further monomer molecules (M) to grow to a higher polymer molecule ((TPP)Al-(M)_n-X). By virtue of the 'living' character of the polymerization, block copolymers ((TPP)Al-(M')_m-(M)_n-X) with narrow MWDs can be synthesized by addition of different monomers (M') to (TPP)Al-(M)_n-X.⁵⁻¹⁰

The polymerization with **1** is strongly affected by the structure of the porphyrin ligand of **1**, since the nucleophilic growing species always carries a (porphinato)aluminium, derived from **1**, as the counter species.¹¹⁻¹⁴ In the course of this study, we have discovered that bulky Lewis acids such as **2** dramatically accelerate the polymerization without spoiling the living character, and the polymerization is called 'Lewis acid-assisted high-speed living

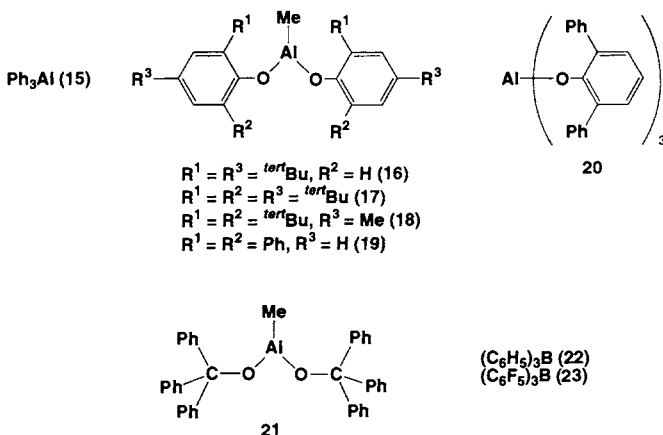
Scheme 14.2. Monomers of living polymerization with metallorhorphyrins³.

anionic polymerization'.¹⁵ Such bulky organoaluminium phenolates have been pioneered by Yamamoto *et al.* in organic syntheses.^{16–17}

2. Principle, scope, and limitations of 'Lewis acid-assisted high-speed living anionic polymerization'

Detailed studies on the above-mentioned accelerated polymerization have shown that nucleophiles (initiators or active polymer terminals) attached to aluminium porphyrins and bulky Lewis acids do not directly interact (react) with each other but can coexist due to a steric repulsion between them.^{16,17} Therefore, the Lewis acids can retain their inherently high ability to coordinate with and activate Lewis basic monomers for nucleophilic attack. This is the basic principle of the 'Lewis acid-assisted high-speed living anionic polymerization'.¹⁸

As shown in Scheme 14.3, representative Lewis acids for the high-speed living anionic polymerization involve methylaluminium diphenoxides having *ortho*-substituents at the phenyl rings.^{18,19} In sharp contrast, with simple trialkylaluminium compounds and methylaluminium diphenoxides without *ortho*-substituents at the phenyl rings, the polymerization is terminated before completion due to undesired direct reactions between the nucleophilic growing species and the Lewis acids, resulting in the formation of a polymer with a broad MWD. An exceptional case is the polymerization with 'hindered' aluminium tetraphenylporphyrins as initiators, where simple trialkylaluminium compounds without any steric protection are usable as accelerators. For



Scheme 14.3. Lewis acids as monomer activators for high-speed living anionic polymerization with aluminium porphyrins as nucleophilic initiators¹⁸.

example, the polymerization of methyl methacrylate initiated with methylaluminium tetramesitylporphyrin is accelerated by the addition of tri-isobutylaluminium, and proceeds to attain 100% monomer conversion, affording a polymer with a narrow MWD.²⁰ Triphenylaluminium is also usable, but the activity is lower than the hindered aluminium phenoxides. Trialkylboron compounds, in sharp contrast with trialkylaluminums, do not terminate nor accelerate the polymerization. On the other hand, triarylborons such as triphenylboron and tris(pentafluorophenyl)boron serve as accelerators for the living polymerization.²¹

In addition to the acceleration of anionic polymerization, combined use of bulky nucleophiles and Lewis acids also allows controlled polymerization of monomers with cationic polymerizability: Although cyclic ethers such as oxiranes are polymerizable both anionically and cationically, oxetanes are only cationically polymerizable due to their strong basicity, and no example had been reported for the controlled synthesis of polyoxetanes because of inherently high potentials of cationic processes for side reactions. In contrast, by using an aluminium porphyrin as nucleophilic initiator in conjunction with a bulky Lewis acid, the polymerization of non-substituted oxetane proceeds in a living anionic fashion to give a polymer with a narrow MWD.²² A sequential high-speed living anionic polymerization of methyl methacrylate and oxetane with this amphiphilic initiating system results in the formation of a narrow MWD block copolymer. Without this method, such a block copolymer could not be available, since methacrylic esters are polymerizable anionically and radically but not cationically.

In place of aluminium porphyrins, Schiff base and phthalocyanine complexes of aluminium can be used in conjunction with bulky Lewis acids for the controlled polymerization of cyclic ethers.²³ Later, this method has been

extended to much simpler systems composed of quaternary onium salts and bulky Lewis acids, which allow high-speed living anionic polymerization of oxiranes and oxetanes.²⁴ In this case, bulky ate complexes are formed from onium salts and Lewis acids, which attack activated monomers through coordination with excess Lewis acids. Alcohol/bulky Lewis acid systems are also effective for the polymerization of lactones and cyclic carbonates.²⁵ More recently, organolithium compounds coupled with bulky Lewis acids have been found to initiate the living stereospecific polymerization of methacrylates at a low temperature to give narrow MWD polymers rich in heterotactic sequence.²⁶

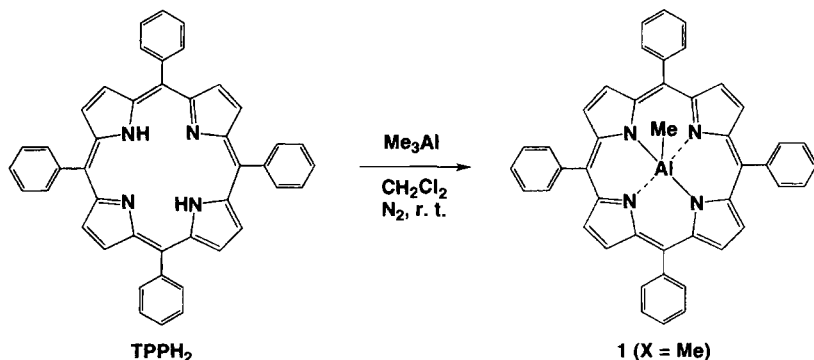
Protocol 1.

Accelerated synthesis of a narrow MWD poly(methyl methacrylate). Polymerization of methyl methacrylate (12, = Me) initiated with methylaluminium 5, 10, 15, 20-tetraphenylporphine (1, X = Me) in the presence of methylaluminium bis(2, 6-di-*tert*-butyl-4-methylphenoxy) (18).²⁸

The procedure consists of three steps; (1) preparation of methylaluminium 5, 10, 15, 20-tetraphenylporphine (TPP)AlMe (1, X = Me), (2) preparation of methylaluminium bis(2, 6-di-*tert*-butyl-4-methylphenoxy) (18) and (3) polymerization of methyl methacrylate (MMA; 12 R x5 Me).

Step 1. Preparation of methylaluminium 5, 10, 15, 20-tetraphenylporphine ((TPP)AlMe) (1, X = Me) (Scheme 14.4)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.4

Protocol 1. Continued**Equipment**

- Magnetic stirrer
- One-necked, round-bottomed flask (100 mL)
- Three-way stopcock
- Teflon-coated magnetic stirring bar (1 × 0.4 cm)
- All-glass syringes (0.25 and 50 mL) with a needle-lock luer
- Needles
- Source of dry nitrogen (preferably from a nitrogen line)

Materials

- 5, 10, 15, 20-tetraphenylporphine (TPPH₂)^a (FW 614.7), 0.615 g, 1 mmol **light-sensitive**
- Trimethylaluminium^b (FW 72.1), 0.096 mL, 1 mmol **pyrophoric, air- and moisture-sensitive**
- Dichloromethane (FW 84.9),^c 40 mL **irritant, toxic**

1. Clean all glassware, syringes, needles, and stirring bar and dry for at least 5 h in a 110°C electric oven before use.
2. Put purified TPPH₂ (0.614 g, 1 mmol) and a Teflon-coated magnetic stirring bar in a one-necked, round-bottomed flask, and equip the neck of the flask with a three-way stopcock using a Demnum grease (Daikin).
3. Support the apparatus using a clamp and a stand with a heavy base, and connect the apparatus to a vacuum/nitrogen line via the three-way stopcock.
4. Dry the above apparatus with a hair dryer under vacuum (10⁻² mm Hg) for 1 h, then back-fill the apparatus with nitrogen. Repeat to a total of three times. Do not use an electric heat gun for drying, since overheating leads to decomposition of TPPH₂.
5. Charge the apparatus containing TPPH₂ with distilled dichloromethane (40 mL), using a syringe through the three-way stopcock in a nitrogen stream, to give a purple suspension of TPPH₂.
6. Support a flask, attached to a three-way stopcock, containing distilled trimethylaluminium under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
7. Fill a syringe with trimethylaluminium (0.096 mL, 1 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent dropwise to the reaction apparatus containing the suspension of TPPH₂ in a nitrogen stream at room temperature. The reaction mixture evolves methane gas and gradually turns homogeneous in 2 h with a colour change from purple to greenish purple, characteristic of methylaluminium 5, 10, 15, 20-tetraphenylporphine ((TPP)AlMe; **1**, X = Me), which displays the appropriate ¹H NMR (in CDCl₃). (TPP)AlMe is sensitive to oxygen and moisture, and easily decomposes to the corresponding alkoxide on exposure to air.

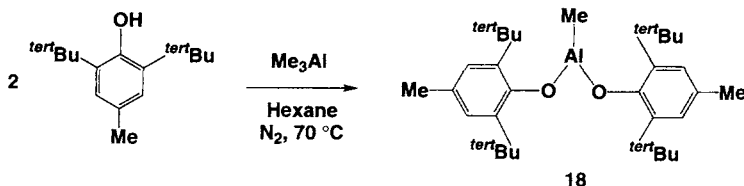
^a Synthesize TPPH₂ by condensation of benzaldehyde (0.8 mol) and pyrrole (0.8 mol) in refluxing propionic acid (2.5 L) for 0.5 h under air.²⁷ Isolate the crystalline precipitate and recrystallize it from chloroform/methanol after washing with hot water (~20% yield). Commercial TPPH₂ (Aldrich) can also be used but must be purified in a similar manner as described above.

^b Distil trimethylaluminium fractionally under reduced pressure in a nitrogen atmosphere.

^c Wash commercial dichloromethane subsequently with concentrated sulfuric acid/water/aqueous NaHCO₃/water in order to remove stabilization agents, and distill fractionally after refluxing over calcium hydride under nitrogen.

Step 2. Preparation of methylaluminium bis(2, 6-di-*tert*-butyl-4-methylphenolate) (**18**) (Scheme 14.5)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.5

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (50 mL)
- Three-way stopcock
- Teflon-coated magnetic stirring bar (1 × 0.4 cm)
- All-glass syringes (1 and 20 mL) with a needle-lock luer
- Needles
- Source of dry nitrogen (preferably from a nitrogen line)

Materials

- 2, 6-Di-*tert*-butyl-4-methylphenol^a (FW 220.4), 2.75 g, 12.5 mmol **irritant**
- Trimethylaluminium^b (FW 72.1), 0.6 mL, 6.25 mmol **pyrophoric, air- and moisture-sensitive**
- Hexane (FW 86.2)^c 10 mL **flammable, irritant**

1. Clean all glassware, syringes, needles, and stirring bar and dry for at least 5 h in a 110°C electric oven before use.
2. Put recrystallized 2, 6-di-*tert*-butyl-4-methylphenol (2.75 g, 12.5 mmol) and a Teflon-coated magnetic stirring bar in a one-necked, round-bottomed flask, and equip the neck of the flask with a three-way stopcock using a Demnum grease.
3. Support the apparatus using a clamp and a stand with a heavy base, and connect the apparatus to a vacuum/nitrogen line via the three-way stopcock.
4. Dry the above apparatus under vacuum (10^{-2} mm Hg) for 1 h, then back-fill the apparatus with nitrogen. Repeat to a total of three times. Do not heat the apparatus, so as to avoid sublimation of 2, 6-di-*tert*-butyl-4-methylphenol.
5. Charge the apparatus containing 2, 6-di-*tert*-butyl-4-methylphenol with distilled hexane (10 mL), using a syringe through the three-way stopcock in a nitrogen stream, to dissolve the phenol.
6. Support a flask, attached to a three-way stopcock, containing distilled trimethylaluminium under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.

Protocol 1. Continued

7. Fill a syringe with trimethylaluminum (0.6 mL, 6.25 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent in the syringe, through the three-way stopcock in a nitrogen stream, dropwise at 0°C to the reaction apparatus containing the hexane solution of 2, 6-di-*tert*-butyl-4-methylphenol.
8. Stir the content of the above apparatus magnetically at room temperature under nitrogen. The reaction mixture evolves methane gas and produces white precipitates in the initial 5–10 min. Stir the suspension for an additional 2 h at room temperature under nitrogen.
9. Warm the suspension at 70°C under nitrogen until the precipitate completely dissolves. Allow the clear solution to stand overnight at room temperature, to give white crystals.
10. Remove a supernatant, liquid phase with a syringe from the above apparatus in a nitrogen stream, and add distilled hexane (5 mL) by syringe to the residue and stir the mixture for a while. Repeat this procedure to a total of three times in order to wash the crystals.
11. Dry the crystals under reduced pressure (10^{-2} mm Hg) for 1 h at room temperature to produce methylaluminum bis(2, 6-di-*tert*-butyl-4-methylphenolate) (**18**) in 64% yield (1.9 g, 4.0 mmol).

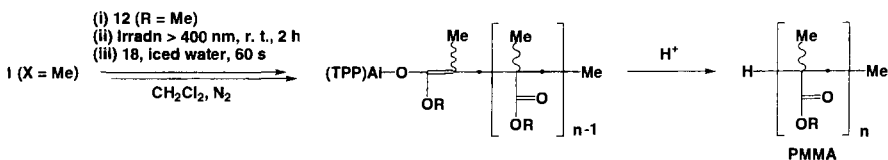
^a Recrystallize commercial 2, 6-di-*tert*-butyl-4-methylphenol from hexane.

^b Distil trimethylaluminum fractionally under reduced pressure in a nitrogen atmosphere.

^c Distil hexane after refluxing over sodium wire under nitrogen.

Step 3. Polymerization of methyl methacrylate (MMA) (Scheme 14.6)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.6

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (100 mL)
- Three-way stopcock
- Beaker (500 mL)
- Teflon-coated magnetic stirring bars (1 × 0.4 cm, 2 × 0.6 cm)
- Needles
- All-glass syringes (2, 20, and 30 mL) with a needle-lock luer
- Source of dry nitrogen (preferably from a nitrogen line)
- 300-W Xenon arc lamp equipped with thermal- and UV-cutoff filters
- Iced water bath

Materials

- (5, 10, 15, 20-Tetraphenylporphinato)aluminium methyl ((TPP)AlMe; **1**, X = Me) (FW 654.7) 0.655 g, 1 mmol in dichloromethane (40 mL) (prepared in Step 1) **air- and moisture-sensitive, light-sensitive**
- Methylaluminium bis(2, 6-di-*tert*-butyl-4-methylphenoxide) (**18**) (FW 480.4), 1.44 g, 3.0 mmol in dichloromethane (10 mL) (prepared in step 2) **air- and moisture-sensitive**
- Methyl methacrylate (MMA; **12**, R = Me)^a (FW 100.1), 20 mL, 200 mmol **flammable, corrosive**
- Methanol (FW 32.0), 500 mL **flammable, toxic**
- Benzene (FW 78.1) 200 mL **flammable, toxic**

1. Support the apparatus containing a dichloromethane solution (40 ml) of **1** (X = Me) (0.655 g, 1 mmol; Step 1, using a clamp and a stand with a heavy base, and connect the apparatus to a nitrogen line via the three-way stopcock.
2. Support a flask, attached to a three-way stopcock, containing distilled MMA (**12** R = Me) under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
3. Fill a syringe with **12** (R = Me) (20 mL, 200 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent in the syringe, through the three-way stopcock in a nitrogen stream, dropwise to the reaction apparatus containing the dichloromethane solution of **1** (X = Me).
4. Stir the contents of the reaction apparatus magnetically at room temperature, and expose for 2 h to a 300-W xenon arc light through thermal- and UV-cutoff filters from a distance of 10 cm. During this period, the solution turns from greenish purple to dark reddish purple, characteristic of an enolatoaluminium porphyrin.
5. Stop the irradiation, and put the reaction apparatus in an iced water bath.
6. Support a flask, attached to a three-way stopcock, containing a dichloromethane solution (0.3 M) of **18** (Step 2) under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
7. Fill a syringe with the dichloromethane solution (10 mL) of **18** (1.4 g, 3 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the solution in the syringe dropwise to the reaction apparatus containing the above photoirradiated reaction mixture of **1** (X = Me) and **12** (R = Me) while stirring magnetically under nitrogen. The polymerization proceeds with a considerable evolution of heat, and reaches 100% monomer conversion within a few seconds.
8. After 30 s, add methanol (1 mL, 24.7 mmol) to the reaction apparatus containing the polymerization mixture, and pour the contents into a large volume of methanol (500 mL) with vigorous stirring magnetically, to give a precipitate.

Protocol 1. Continued

9. Collect the precipitate, dissolve it in benzene (200 mL), and freeze-dry to give a poly(methyl methacrylate) ($M_n = 21500$, $M_w/M_n = 1.10$) in 90% yield (18 g) based on the charged monomer.

^a Distil **12** (R = Me) after stirring with calcium hydride or more vigorously with trimethylaluminum under nitrogen.

By changing the mole ratio of **12** (R = Me) to **1** (X = Me), the molecular weight of the polymer can be controlled over a wide range up to 10^6 , retaining the narrow MWD ($M_w/M_n = 1.05\text{--}1.2$).²⁸ The present procedure is also applicable to the living anionic polymerization of other methacrylates (**12**; R = ethyl, isopropyl, n-butyl, isobutyl, benzyl, and dodecyl,²⁹ and methacrylonitrile (**13**).³⁰

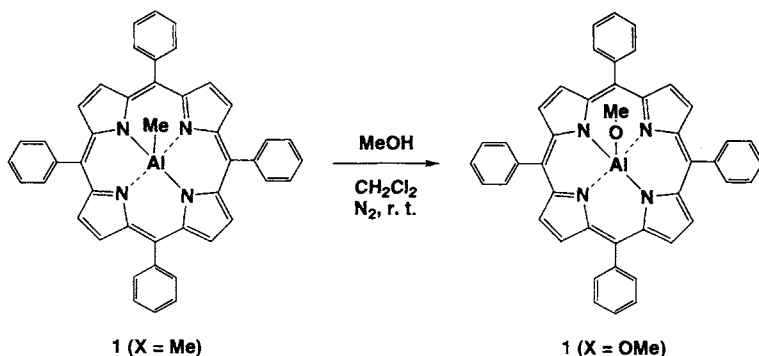
Protocol 2.

Accelerated synthesis of a narrow MWD poly(δ -valerolactone).
Polymerization of δ -valerolactone (5**) initiated with (5, 10, 15, 20-tetraphenylporphinato)aluminium methoxide ((TPP)AlOMe; **1**, X = OMe) in the presence of methylaluminium bis(2, 6-diphenylphenoxide) (**19**)³¹**

The procedure consists of three steps; (1) preparation of (5, 10, 15, 20-tetraphenylporphinato)aluminium methoxide ((TPP)AlOMe; **1**, X = OMe), (2) preparation of methylaluminium bis(2, 6-diphenylphenoxide) (**19**) and (3) polymerization of δ -valerolactone (**5**).

Step 1. Preparation of (5, 10, 15, 20-tetraphenylporphinato)aluminium methoxide ((TPP)AlOMe; **1**, X = OMe) (Scheme 14.7)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.7

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (50 ml)
- Three-way stopcock
- Teflon-coated magnetic stirring bar (1 × 0.4 cm)
- All-glass syringes (10 and 30 mL) with a needle-lock luer
- Needles
- Source of dry nitrogen (Preferably from a nitrogen line)

Materials

- Methylaluminium 5, 10, 15, 20-tetraphenylporphine ((TPP)AlMe; **1**, X = Me)^a (FW 654.7), 0.065 g, 0.1 mmol in dichloromethane (4 mL) (for preparation, see Protocol 1, Step 1)

air- and moisture-sensitive, light-sensitive
flammable, toxic

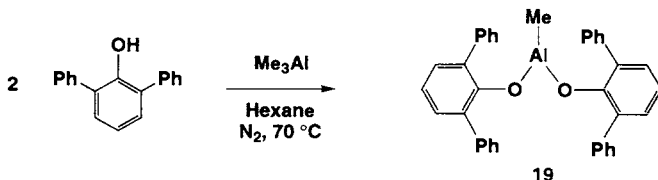
- Methanol (FW 32.0),^a 1 mL, 24.6 mmol

1. Support the reaction apparatus, containing a dichloromethane solution (4 mL) of **1** (X = Me) (0.065 g, 0.1 mmol), under nitrogen (see Protocol 1, Step 1), using a clamp and a stand with a heavy base, and connect the apparatus to a vacuum/nitrogen line via the three-way stopcock.
2. Support a flask, attached to a three-way stopcock, containing distilled methanol under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
3. Fill a syringe with methanol (1 mL, 24.6 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent in the syringe, through the three-way stopcock in a nitrogen stream, to the reaction apparatus containing the dichloromethane solution of **1** (X = Me) at room temperature.
4. Stir the contents of the reaction apparatus magnetically for 15 h at room temperature under nitrogen. The solution gradually turns from greenish purple to bright reddish purple, characteristic of a (porphinato)aluminium alkoxide.
5. Connect the reaction apparatus to a vacuum line, and remove volatile fractions from the contents under reduced pressure (10^{-2} mm Hg), and back-fill the apparatus with nitrogen, leaving (TPP)AlOMe (**1**, X = OMe) as a reddish purple powder, which displays the appropriate ¹H NMR (in CDCl₃).

^a Distil methanol after refluxing over magnesium ribbon under nitrogen.

Step 2. Preparation of methylaluminium bis(2, 6-diphenylphenoxide) (**19**) (Scheme 14.8)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.8

Protocol 2. Continued

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (5 mL)
- Three-way stopcock
- Teflon-coated magnetic stirring bar (1 × 0.4 cm)
- All-glass syringes (0.5 and 20 mL) with a needle-lock luer
- Needles
- Source of dry nitrogen (preferably from a nitrogen line)

Materials

- Trimethylaluminum^a (FW 72.09), 0.29 mL, 3 mmol **pyrophoric, air- and moisture-sensitive**
- 2, 6-Diphenylphenol^b (FW 246.3), 1.5 g, 6 mmol **irritant**
- Dichloromethane (FW 84.0,^c 10 mL **irritant, toxic**)

1. Clean all glassware, syringes, needles, and stirring bar and dry for at least 5 h in a 110°C electric oven before use.
2. Put recrystallized 2, 6-diphenylphenol (1.5 g, 6 mmol) and a Teflon-coated magnetic stirring bar in a one-necked, round-bottomed flask, and equip the neck of the flask with a three-way stopcock using a Demnum grease (Daikin).
3. Support the apparatus using a clamp and a stand with a heavy base, and connect the apparatus to a vacuum/nitrogen line via the three-way stopcock.
4. Dry the above apparatus under vacuum (10^{-2} mm Hg) for 1 h, then back-fill the apparatus with nitrogen. Repeat to a total of three times. Do not heat the apparatus, so as to avoid sublimation of 2, 6-diphenylphenol.
5. Using a syringe, add dichloromethane (10 mL) to the apparatus containing 2, 6-diphenylphenol through the three-way stopcock in a nitrogen stream, to dissolve the phenol.
6. Support a flask, attached to a three-way stopcock, containing distilled trimethylaluminium under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
7. Fill a syringe with trimethylaluminium (0.29 mL, 3 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent in the syringe, through the three-way stopcock in a nitrogen stream, dropwise to the reaction apparatus containing the dichloromethane solution of 2, 6-diphenylphenol at 0°C.
8. Stir the contents of the apparatus at room temperature under nitrogen. The reaction mixture evolves methane gas during the initial 5–10 min. Stir the solution for an additional 2 h at room temperature under nitrogen. Use the resulting pale-yellow solution containing **19** for the subsequent polymerization (Step 3).

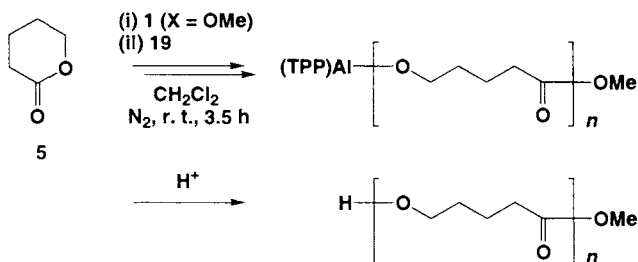
^a Distil trimethylaluminium fractionally under reduced pressure in a nitrogen atmosphere.

^b Recrystallize commercial 2, 6-diphenylphenol from hexane.

^c Wash commercial dichloromethane successively with concentrated sulfuric acid/water/aqueous NaHCO₃/water in order to remove stabilization agents, and fractionally distil after refluxing over calcium hydride under nitrogen.

Step 3. Polymerization of δ -valerolactone (**5**) (Scheme 14.9)

Caution! Carry out all procedures in a well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.9

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (100 mL)
- Three-way stopcock
- Teflon-coated magnetic stirring bars (1 × 0.4 cm, 2 × 0.6 cm)
- All-glass syringes (2, 5, and 20 mL) with a needle-lock luer
- Needles
- Source of dry nitrogen (preferably from a nitrogen line)

Materials

- (5, 10, 15, 20-Tetraphenylporphinato)aluminium methoxide ((TPP)AlOMe; **1**, X = OMe)^a (FW 670.7), 0.067 g, 0.1 mmol (prepared in Step 1)
- Methylaluminium bis(2, 6-diphenylphenoxide) (**19**) (FW 532.6) 0.16 g, 0.3 mmol in dichloromethane (10 mL) (prepared in Step 2)
- δ -Valerolactone (**5**)^a (FW 100.1), 1.86 mL, 20 mmol
- Dichloroethane^b (FW 98.96), 4 mL, 50 mmol
- methanol (FW 32.0), 500 mL
- Benzene (FW 78.1), 200 mL

moisture-sensitive, light-sensitive

air- and moisture-sensitive
irritant
flammable, toxic
flammable, toxic
flammable, toxic

1. Support the apparatus containing **1** (X = OMe) (0.067 g, 0.1 mmol; Step 1) under nitrogen, using a clamp and a stand with a heavy base, and connect the apparatus to a nitrogen line via the three-way stopcock.
2. Using a syringe add distilled dichloroethane (4 mL) through the three-way stopcock in a nitrogen stream, to give a clear, bright reddish purple solution.
3. Support a flask, attached to a three-way stopcock, containing distilled **5** under nitrogen, using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
4. Fill a syringe with **5** (1.86 mL, 20 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the reagent in the syringe, through the three-way stopcock in a nitrogen stream, dropwise to the reaction apparatus containing the dichloroethane solution of **1** (X = OMe) upon stirring magnetically at room temperature.

Protocol 2. Continued

- Support a flask, attached to a three-way stopcock, containing a dichloromethane solution (0.3 M) of **19** under nitrogen (prepared in Step 2), using a clamp and a stand with a heavy base, and connect the apparatus to a nitrogen line via the three-way stopcock.
- Fill a syringe with the dichloromethane solution (1 mL) of **19** (0.16 g, 0.3 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the solution in the syringe, through the three-way stopcock in a nitrogen stream, to the reaction apparatus containing the mixture of **1** ($X = \text{OMe}$) and **5** in dichloroethane upon stirring magnetically at room temperature.
- Stir the mixture in the reaction apparatus magnetically for 3.5 h at room temperature under nitrogen. The polymerization reaches 74% monomer conversion to give a viscous, partially solidified solution.
- Pour the polymerization mixture in a large volume of methanol (500 mL) with vigorous stirring, to give a precipitate.
- Collect the precipitate, dissolve it in benzene (200 mL), and freeze-dry, to give a poly(δ -valerolactone) ($M_n = 24000$, $M_w/M_n = 1.13$) in 67% yield (1.3 g) based on the charged monomer.

^a Distil **5** after stirring with calcium hydride under nitrogen.

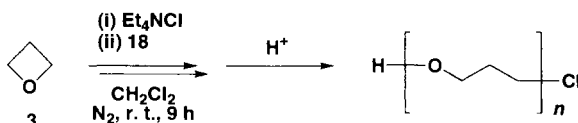
^b Wash commercial dichloroethane successively with concentrated sulfuric acid/water/aqueous NaHCO_3 /water in order to remove stabilization agents, and fractionally distil after refluxing over calcium hydride under nitrogen.

By changing the mole ratio of **5** to **1** ($X = \text{OMe}$), the molecular weight of the polymer can be controlled over a wide range up to 10^5 , retaining the narrow MWD ($M_w/M_n < 1.2$). Use of **18** as accelerator results in broadening of MWD at a high conversion ($>70\%$), due to undesired transesterification of the produced polymer caused by **18**.³¹

Protocol 3.

Accelerated synthesis of a narrow MWD polyoxetane. Polymerization of oxetane (3**) initiated with tetraethylammonium chloride (Et_4NCl) in the presence of methylaluminum bis(2, 6-di-*tert*-butyl-4-methylphenolate) (**18**).²⁴ (Scheme 14.10)**

Caution! Carry out all procedures in well-ventilated hood, and wear disposable vinyl or latex gloves and chemical-resistant safety goggles.



Scheme 14.9

14: Lewis acid-assisted anionic polymerizations

Equipment

- Magnetic stirrer
- One-necked, round-bottomed flask (50 mL)
- Three-way stopcock
- Teflon-coated magnetic stirring bar (1 × 0.4 cm)
- Needles
- All-glass syringes (1, 2, and 3 mL) with a needle-lock luer
- Source of dry nitrogen (preferably from a nitrogen line)
- Glass column for chromatography (50 × 4 cm)

Materials

- Tetraethylammonium chloride (Et_4NCl)^a (FW 165.7), 0.0166 g, 0.1 mmol **hygroscopic, irritant**
- Methylaluminum bis(2, 6-di-*tert*-butyl-4-methylphenolate) (**18**)^b (FW 480.7), (0.14 g, 0.3 mmol) in dichloromethane solution (0.3 M, 1 mL) (for preparation, see Protocol 1, Step 2) **air- and moisture-sensitive**
- Dichloromethane,^b 2 mL **irritant, toxic**
- Oxetane (**3**) (FW 58.1),^c 0.65 mL, 10 mmol **flammable**
- Silica gel (C300) 150 g

1. Clean all glassware, syringes, needles, and stirring bar and dry for at least 5 h in a 110°C electric oven before use.
2. Put recrystallized tetraethylammonium chloride (Et_4NCl) (0.0166 g) and a Teflon-coated magnetic stirring bar in a one-necked, round-bottomed flask, and equip the neck of the flask with a three-way stopcock using a Demnum grease (Daikin).
3. Support the above apparatus using a clamp and a stand with a heavy base, and dry the content under vacuum (10^{-2} mm Hg) for 1 h, then back-fill the apparatus with nitrogen. Repeat to a total of three times. Do not heat the apparatus, so as to avoid decomposition of Et_4NCl .
4. Using a syringe add dichloromethane (2 mL) through the three-way stopcock in a nitrogen stream, to dissolve Et_4NCl .
5. Support a flask, attached to a three-way stopcock, containing a dichloromethane solution of **18** (0.3 M) under nitrogen (Protocol 1, Step 2), using a clamp and a stand with a heavy base, and connect the flask to a nitrogen line via the three-way stopcock.
6. Fill a syringe with the dichloromethane solution (1 mL) of **18** (0.14 g, 0.3 mmol) from the flask through the three-way stopcock in a nitrogen stream, and add the solution in the syringe, through the three-way stopcock in a nitrogen stream, dropwise to the reaction apparatus containing the dichloromethane solution of Et_4NCl while stirring magnetically at room temperature.
7. Likewise, add **3** (0.65 mL, 10 mmol) to the above reaction apparatus, and stir the contents for 9 h at room temperature. The polymerization reaches 100% monomer conversion, to give a viscous solution.
8. Evaporate the polymerization mixture to dryness with a rotary evaporator. Separate the residue by silica gel column chromatography with ethyl acetate–hexane (4:1) as eluant. Evaporate the fraction containing polymeric

Protocol 3. Continued

products with a rotary evaporator, to leave a polyoxetane as a colourless viscous liquid ($M_n = 5200$, $M_w/M_n = 1.12$) in 80% yield (0.46 g) based on the monomer charged.

^a Recrystallize commercial Et_4NCl from acetone–hexane.

^b Wash commercial dichloromethane successively with concentrate sulfuric acid/water/aqueous NaHCO_3 /water in order to remove stabilization agents, and distil fractionally after refluxing over calcium hydride under nitrogen.

^c Distil commercial **3** over sodium wire under reduced pressure in a nitrogen stream at room temperature.

By changing the mole ratio of **3** to Et^4NCl , the molecular weight of the polymer can be controlled up to 20 000, retaining the narrow MWD ($M_w/M_n = 1.09$ – 1.15).²⁴ No polymerization takes place when the mole ratio **18**: Et_4NCl is less than unity.

References

1. Szwarc, M. *Nature* **1956**, *178*, 1168–1170.
2. Szwarc, M.; Levy, M.; Milkovich, R. *J. Am. Chem. Soc.* **1956**, *78*, 2656–2657.
3. Aida, T.; Inoue, S. *Acc. Chem. Res.* **1996**, *29*, 39–48.
4. Aida, T. *Prog. Polym. Sci.* **1994**, *19*, 469–528.
5. Aida, T.; Inoue, S. *Macromolecules* **1981**, *14*, 1162–1166.
6. Yasuda, T.; Aida, T.; Inoue, S. *Macromolecules* **1984**, *17*, 2217–2222.
7. Aida, T.; Sanuki, T.; Inoue, S. *Macromolecules* **1985**, *18*, 1049–1055.
8. Aida, T.; Ishikawa, M.; Inoue, S. *Macromolecules* **1986**, *19*, 8–11.
9. Kuroki, M.; Nashimoto, S.; Aida, Y.; Inoue, S. *Macromolecules* **1988**, *21*, 3114–3115.
10. Hosokawa, Y.; Kuroki, M.; Aida, T.; Inoue, S. *Macromolecules* **1991**, *24*, 824–829.
11. Aida, T.; Inoue, S. *J. Am. Chem. Soc.*, **1985**, *107*, 1358–1364.
12. Kuroki, M.; Aida, T.; Inoue, S. *J. Am. Chem. Soc.*, **1987**, *109*, 4737–4738.
13. Sugimoto, H.; Aida, T.; Inoue, S. *Macromolecules* **1990**, *23*, 2869–2875.
14. Watanabe, Y.; Aida, T.; Inoue, S. *Macromolecules* **1991**, *24*, 3970–3972.
15. Kuroki, M.; Watanabe, T.; Aida, T.; Inoue, S. *J. Am. Chem. Soc.* **1990**, *112*, 5639–5640.
16. Yamamoto, H.; Maruoka, K.; Furuta, K.; Naruse, Y. *Pure Appl. Chem.* **1988**, *61*, 419–422.
17. Yamamoto, H.; Maruoka, K. *J. Syn. Org. Soc. Jpn* **1993**, *51*, 1074–1086.
18. Inoue, S.; Aids, T. *CHEMTECH*, **1994**, *24*, 28–35.
19. Aida, T.; Metalloporphyrin catalysis. In *Catalysis in precision polymerization* (ed S. Kobayashi). Wiley, New York, pp 310–322.
20. Sugimoto, H.; Aida, T.; Inoue, S. *Macromolecules* **1994**, *27*, 3672–3674.
21. Sugimoto, H.; Aida, T.; Inoue, S. *Macromolecules* **1993**, *26*, 4751–4755.
22. Takeuchi, D.; Watanabe, Y.; Aida, T.; Inoue, S. *Macromolecules* **1995**, *28*, 651–652.
23. Sugimoto, H.; Kawamura, C.; Kuroki, M.; Aida, T.; Inoue, S. *Macromolecules* **1994**, *27*, 2013–2018.

24. Takeuchi, D.; Aida, T. *Macromolecules* **1996**, 29, 8096–8100.
25. Akatsuka, M.; Aida, T.; Inoue, S. *Macromolecules* **1995**, 28, 1320–1322.
26. Kitayama, T.; Zhang, Y.; Hatada, K. *Polym. J.* **1994**, 26, 868–872.
27. Adler, A. D. *J. Org. Chem.* **1967**, 32, 476–477.
28. Sugimoto, H.; Kuroki, M.; Watanabe, T.; Kawamura, C.; Aida, T.; Inoue, S. *Macromolecules* **1993**, 26, 3403–3410.
29. Sugimoto, H.; Saika, M.; Hosokawa, Y.; Aida, T.; Inoue, S. *Macromolecules* **1996**, 29, 3359–3369.
30. Adachi, T.; Sugimoto, H.; Aida, T.; Inoue, S. *Macromolecules* **1992**, 25, 2280–2281.
31. Isoda, M.; Sugimoto, H.; Aida, T.; Inoue, S. *Macromolecules* **1997**, 30, 57–62.

Index

Note: Bold type denotes chapter extent

- acetalization 160, 163–4, 189–90
- acetone cyanohydrin 220
- acetophenone 12, 240
- acetophenone-rhenium complexes 226–7
- 3-acetyl-1,3-oxazolidin-2-one 196
- acetylation of alcohol 190–2
- L*-acetyl menthol 191–2
- acid anhydrides 190–2, 214
- acrolein derivatives, *see* bromoacrolein; methacrolein
- acrylic acid 40
- 3-acryloxazolidine-2-one 68–9
- 3-acryloyl-1,3-oxazolin-2-one 232–4
- (acycloxy)boranes, *see* chiral (acycloxy)boranes (CAB)
- acycloxy naphthalenes 186
- acylation, *see* Friedel–Crafts reaction
- acylgermane 22
- acyl halides 186
- acylimides 230, 231, 233, 237
- acylsilanes 22, 23
- aglycones 122
- alcohols
 - acetylation and esterification 190–2
 - allylation of arenes 186–7
 - asymmetric epoxidation of allylic 26
 - cyclopropanation of allylic 78–9, 86–7
 - synthesis of equatorial 6–10
 - synthesis of homoallyl 53, 163, 192
- aldehydes
 - acetalization 160, 163–4, 189–90
 - and α -chloroketones 217
 - and alkenylzirconocene chloride 181–2
 - and allylsilanes 100
 - and allylstannanes 100–1
 - and alkylzincs 78, 79–83, 96, 97–9
 - cyanation/cyanoisilylation 129–30, 151–3, 208
 - ether formation 164–7
 - formation from epoxy substrates 24–8
 - hydrophosphonylation 217
 - hydrosilylation of aromatic 35
 - and isocyanides 169–73
 - and vinylic ethers 215
 - see also in particular* aldol and aldol-type reactions; alkylation; allylation; Diels–Alder reactions; *named aldehydes*
- aldol and aldol-type reactions
 - and boron reagents 32, 51–4, 55–7
 - imine reactions 33–4, 58–61
 - and cerium enolates 204, 207–8
 - and copper(II) complexes 240–2
 - and europium(III) complexes 208, 211–12
 - of isocyanides using silver/gold complexes 169–73
 - and lanthanum(III) reagents 217
 - and lithium perchlorate 65, 66, 71
 - and palladium catalysts 240
 - and rhodium complexes 238, 239
 - and ruthenium complexes 238, 239–40
 - and scandium(III) triflate 194–6
 - and silicon(IV) reagents 161–3
 - and tin Lewis acids 138–51
 - and titanium complexes 113–18
 - and ytterbium(III) reagents 214, 216
- aldols: selective dehydration 35–6, 37
- alkali and alkaline earth metal salts **65–75**
 - aldol additions and allylations 71–4
 - cycloaddition reactions 65–70
- alkaloids 140, 199, 200
- alkenes
 - 1,3-dipolar cycloaddition 69–70
 - hydrosilylation 199
- alkenylzirconocene chloride 180–2
- α -alkoxy aldehydes 67, 71–3, 209, 240
- α -alkylated α -amino acids 169
- alkylations 225–7
 - amphiphilic carbonyl 6–10
 - amphiphilic conjugate 10–12
- alkyldichloroborane 44, 45
- alkyl silyl ethers 160–1, 163
- alkyltitanium complexes 96–101
- alkylzincs 78, 79–83, 96, 97–9
- alkynes, reductive coupling 182–4
- allenes, reductive coupling 182–4
- allylarenes 186–7
- allylation
 - and allyltin/silver(I)-phosphine complex 173–5
 - and aluminium reagents 12–15
 - and boron reagents 51–5
 - and lithium perchlorate 71
 - and magnesium bromide 65, 71–2
 - and silicon(IV) reagents 163–4
 - and tetra-allyltin/scandium(III) triflate 192–4
 - see also* Sakurai–Hosomi reaction
- allylcopper 12
- allylic alcohols, *see* alcohols
- allylic ethers 106–7
- allylsilanes 53, 99–101, 163

- allylsilicates 159
- allylstannanes 71–3, 99–100
- allyltins 55, 173–5, 192–4
- allyltributylstannane 71
- allyl vinyl ethers: Claisen rearrangement 21–3
- aluminium chloride 152
- aluminium porphyrins 246, 247–58
- aluminium reagents **5–29**
 - amphiphilic alkylations 6–15
 - Claisen rearrangement 21–3
 - Diels–Alder reaction
 - asymmetric hetero- 16–20
 - regio- and stereocontrolled 15–16
 - epoxide rearrangement 23–8
 - in polymerization reactions 246, 247–60
 - properties 5
- aluminium thiolate 21
- aluminium tris(2,6-diphenylphenoxide) (ATPH) 10–12
 - in Claisen rearrangement 21, 23
 - derivatives for allylation 13–15
- aluminium tris(2- α -naphthyl-6-phenylphenoxide) (ATNP) 21
- aluminium tris(*R*)-1- α -naphthyl-3-phenylphenoxide) ((*R*)-ATBN) 23
- aluminium tris(*R*)-3-(*p*-fluorophenyl)-1- α -naphthylphenoxide) ((*R*)-ATBN-F) 23
- amidation 3
- amidation catalysts 37–9
- amines, optically active 79
- amino acid-derived Schiff base 130
- β -amino acid esters 33, 59
- amino acids 169–71
 - N*-sulfonyl derivatives of α -amino acids 42
- amino alcohol-derived Schiff base 130
- β -aminoalcohols 78, 79–83
- α -amino aldehydes 169, 209
- β -aminothiols 78
- amphiphilic carbonyl alkylations 6–10
- amphiphilic conjugate alkylations 10–12, 12–15
- anionic polymerization, *see* polymerization
- anisole 186, 187–8
- anti*-Cram selectivity 7, 8
- D*-arabinose 192–4
- arenes 12
 - acylation 185–7
- 2-arlycyclohexanones 153, 154
- arylboron reagents 31, 32–9
- asymmetric amplification 77
- asymmetric catalytic desymmetrization 105–6
- asymmetric cyanation reaction 151–3
- asymmetric Diels–Alder reactions, *see* Diels–Alder reactions
- asymmetric protonation 153–4
- autocatalysis 78
- aza Diels–Alder reaction 57–8
- aziridines 215
- benzaldehyde
 - addition of alkylzincs 79–81, 98
 - aldol reactions 53, 71, 139–40, 143, 145, 146–7, 150–1, 172–3, 195–6, 209, 238, 239–40
 - allylations 54–5, 174–5
 - conjugate alkylations 12
 - Diels–Alder reactions 17, 18–19, 211, 229, 230
- benzene 188–9
- benzene-chrome complexes 78
- N*-benzhydryl imines 59–60
- benzophenone 6
- benzyl alcohol 186–7, 188–9
- benzylidenephénylamine *N*-oxide 69–70
- N*-benzylimines 33, 34
- (*Z*)-2-benzyloxy-1-*S*-ethyl-1-trimethylsiloxyethene 143, 144
- (2*R*,3*R*,4*S*)-2-benzyloxy-4-methyl-5-hexen-3-ol 72–3
- α -(benzyloxy)acetaldehyde 240, 241–2
- α -benzyloxy aldehyde 71, 117
- 2-(benzyloxy)butyraldehyde 67–8
- (2*S*,1'*S*)-2-[(1'-benzyloxy)propyl-6-methyl-2,3-dihydro-4*H*-pyran-4-one 67–8
- (*R*)-BINAL 16–20, 22–3
- BINAP 173–5, 240
- binaphthol, *see* BINOL complexes
- BINOL-bromine derivatives 122, 125
- BINOL-lanthanum complexes 217
- BINOL-tin complexes 153
- BINOL-titanium complexes
 - and aldol reactions 113–18
 - asymmetric catalytic desymmetrization 105–6
 - carbonyl addition reaction 96–101
 - carbonyl-ene reaction 101–5
 - cyanohydrin formation 129–30
 - and Diels–Alder reaction 121–9
 - enantiomer-selective activation of racemic catalysts 108–11
 - ene-cyclization 113
 - kinetic optical resolution 106–7
 - positive NLE of non-racemic catalysts 107–8
- BINOL-ytterbium complexes 214, 215
- bis(3,5-bis(trifluoromethyl)phenyl)borinic acid 31, 35
- bis(imine)-copper(II) complex 233–4
- bis(oxazoline)-copper(II) complex 230, 237, 240, 241
- bis(oxazoline)-iron(III) complex 230, 231, 232–3

- bis(oxazoline)-magnesium(II) complexes 65, 68–70
 bis(oxazoliny)pyridine-copper(II) (pybox) complexes 236, 237–8, 240, 241–2
 bis(pentafluorophenyl)borinic acid 31, 35, 37
 bis(phosphino)ferrocene 169
 (*R*)-(+)-3,3'-bis(triarylsilyl)binaphthol 16, 17
 3,5-bis(trifluoromethyl)phenylboronic acid 31, 37
 2,4-bis(trimethylsiloxy)-1,3-pentadiene 67–8
 boron reagents **31–63**
 arylboron reagents 31, 32–9
 chiral boron reagents 35–61
 boron tribromide 45
 bromoacrolein 42, 43–4, 48–9, 129, 237, 238
 3-bromocamphor 20
 Brønsted acid-assisted chiral Lewis acid (BLA)
 aldol type reaction of imines 59–61
 Diels–Alder reactions 47–51, 58
 butadiene 228
 butanal 165–7
 3-((*E*)-2'-butenyl)-1,3-oxazolidin-2-one 69–70
 4-*tert*-butyl-1-methylcyclohexanol 9–10
 1-butyl-2-propenyl vinyl ether 21
 butylcalcium iodide 13
 4-*tert*-butylcyclohexanone 6, 9
 (*Z*)-2-*t*-butyldimethyl-siloxy-1-*S*-ethyl-1-trimethylsiloxyethene 145–6
tert-butyldimethylsilyl enol ether 153
 butyl-lithium 12, 13, 14, 15
t-butylmagnesium chloride 8
t-butyl methyl fumarate 15–16
 1-*tert*-butylthio-[(trimethylsilyl)oxy]ethene 241–2
 butyrolactone 73–4

 CAB, *see* chiral (acycloxy)boranes
 ϵ -caprolactam: polymerization 221
 carbonyl activation 6, 20, 180–2, 225–7
 carbonyl addition reactions 4, 96–101
 carbonyl alkylations, amphiphilic 6–10
 carbonyl-ene cyclization 112–13
 carbonyl-ene reaction 101–5
 carboxamides 37–9
 carboxylic acids: amidation 37–9
 carnitine 116, 117
 cerium chloride 203–6
 cerium enolates 204, 207–8
 cerium halides 204–5
 cerium(III) reagents 203–8, 219
 cerium iodide 204–5
 chalcone 85–6
 'chiral activators' 109, 110
 chiral (acycloxy)boranes (CAB) 31, 32
 catalysis of enantioselective aldol reactions 52–3
 catalysis of enantioselective Diels–Alder reaction 39–44
 catalysis of Sakurai–Hosomi allylation 54–5
 α -chiral aldehydes 209
 chiral β -aminoalcohols 78, 79–83
 chiral diamines 138, 140–51
 structures 139
 chiral diols 189
 Chiral Lewis Acid-controlled Synthesis (CLAC synthesis) 145, 146–7
 chiralphos 229
 'chiral poisoning' 108
 α -chloroketones 217
 3-cholestanone 7
 (+)-cinchonine trimethylsilyl ether 152
 cinnamaldehyde 12, 13, 14, 15
 cinnamyl alcohol 86–7
 Claisen rearrangement 3, 4, 21–3, 66
 classification of reactions 2, 3, 4
 cobaltocenium ion 229
 compactin 121
 conjugate alkylations 10–12, 12–15
 copper(II) complexes 78
 copper(II)-bis(imine) 233–4
 copper(II)-bisoxazoline 230, 237, 240, 241
 copper(II)-bis(oxazoliny)pyridine (pybox) 236, 237–8, 240, 241–2
 3-crotonoyl-1,3-oxazolidin-2-one 196–9
 crotylation 71
 crotylsilane 163
 cyanation/cyanosilylation reactions 129–30, 151–3, 208, 215
 cyanohydrin formation 129–30, 152
 cyanohydrin trimethylsilyl ethers 152
 cyanotrimethylsilane 129–30
 cyclic ethers 248
 cyclic ketones: alkylation 6–10
 cycloaddition reactions 15, 42, 44, 65–70
 see also Diels–Alder reactions
 cyclohexanecarboxyaldehyde 152
 2-cyclohexene-1-one 206
 cyclohexyloxytrimethylsilane 165–7
 cyclopentadiene
 Diels–Alder reactions
 with 2-bromoacrolein 48–9
 with α,β -unsaturated enones 228–9
 with acylimide 231
 with (*E*)-2-pentenal 50–1
 with fumarates 16
 with methacrolein 39, 40, 44, 45, 234–8
 with methyl acrylate 45, 65, 66–7
 with oxazolidinones 68–9, 196–9, 232
 cyclopentadienyl-manganese complexes 226
 cyclopentadienyl-rhenium complexes 226
 cyclopentadienyl-ruthenium complexes 225–6, 229
 cyclopentadienyl yttrium hydride 199–200

- 2-cyclopentenone 205, 213
 cyclopropanation 78–9, 86–7
- Danishefsky dienes 57, 58, 117, 125, 229
 2-deoxy-D-ribonolactone 211–12
 desymmetrization 105–6
 dialkyl acetals 160
 (*R*)-(+)–3,3'-dialkylbinaphthol 17
 dialkylzincs 77–91, 96, 97–9
 β -diamines 78
 diamines, chiral, *see* chiral diamines
 diamino bisphosphine-gold(I) complex 172–3
 diarylborinic acids 31, 35–7
 diarylmethanes 186–7
 diborane 42
 dibutyltin diacetate 138, 139, 143, 144, 145, 146, 147, 148
 Diels–Alder reactions 2, 4
 and alkali/alkaline earth metal reagents 65–8
 and bulky aluminium reagents 15–16
 chiral 16–20
 and chiral boron reagents 39–51
 imine reactions 57–8
 and chiral titanium complexes 118–29
 and europium(III) complexes 208–11
 inverse electron-demand 130
 and lanthanum(III) reagents 217
 other transition-metal complex catalysts 228–38
 and scandium(III) triflate 196–9
 and ytterbium triflate 214
 dienes, *see in particular* Diels–Alder reactions;
 named dienes
 1,3-dienol derivatives 122–4
 diethylzinc 77, 78, 79–81, 85–7
trans-4,5-dihydro-3-phenyl-4,5,5-trimethyl-2(3*H*)-furanone 73–4
cis-dihydropyrene 17
anti- α,β -dihydroxy thioesters 144
syn- α,β -dihydroxy thioesters 145
 di-isopropoxytitanium (IV) dibromide 95–6
 di-isopropylzinc 78
 1,5-diketones 217
 dimethoxybenzene 186
 1,1-dimethoxybutane 162
 2,4-dimethyl-1-methoxy-3-trimethylsiloxy-1,3-butadiene 20
 (5*R*,6*R*)-3,5-dimethyl-6-phenyldihydropyran-4-one 18–19
 dimethylaluminium 1,1-diphenylethoxide 6
 2,6-dimethylbenzoquinone 65–6
 dimethyl cuprate 227
 dimethylzinc 77, 81–4
 diols 109, 110
 chiral 189
 diol 229
 dioxane 96, 97, 189
 dioxolane 189
 diphenylacetaldehyde 24–6
 disulfonamide-titanate 78, 79, 86–7, 97
- α,β -enals 40, 42, 47, 49
 enamines 226
 ene cyclizations 111–13
 ene reactions 2, 3, 20
 carbonyl-ene reaction 101–8
 intramolecular (cyclizations) 111–13
 pathway in aldol reaction 114–17
 enol derivatives: protonation 153
 enol silyl ethers, *see* silyl enol ethers
 enol trimethylsilyl ether 216
- α,β -enones
 cyanosilylation 215
 from aldol dehydration 36
 reactions with cyclopentadiene 228–9
 reactions with enol silyl ethers 214
 reactions with KSA 208, 209
 reduction of 203, 204
 enoxysilacyclobutane 159
 epilupine 199, 200
 epoxide rearrangement
 activation by zirconocene complex 179
 aluminium reagents 23–8
 boron catalysts/reagents 35
 epoxy silyl ethers 26, 27, 28
 equatorial alcohols 6–10
 esterification of alcohol 190–2
 esters 153, 161
 hydrosilation of 35
 hydroxy esters 52, 101, 102–3
 ethers 160–1, 163, 164–7
 activation by hafnocene complexes 179–80
 kinetic resolution 106–7
 polymerization of cyclic 248–9
 vinyl 21–3, 215
 see also silyl enol ethers
 1-*S*-ethyl-1-trimethylsiloxyethene 138, 140, 141–2
 1-*S*-ethyl-1-trimethylsiloxypropene 139–40, 142, 143, 150–1
 (5*S*)-1-ethyl-2-[(piperidin-1-yl)-methyl]pyrrolidine 144
 ethylene: polymerization 220
S-ethyl propanethioate 144
 europium(III) reagents 208–14, 218
- Felkin–Anh-like transitional state model 117
 ferrocene derivatives 78, 169, 229
 ferrocenium ion 228
 fluoral-ene reaction 104–5
 formaldehyde 216

- Friedel–Crafts reaction 2, 3
 hafnium triflate catalysis 180–1
 scandium triflate catalysis 185–9
 yttrium triflate catalysis 214
 Fries rearrangement 180–1, 186
 fumarates 15–16
 furan 42, 43–4
- GABOB** 116, 117
 gadolinium compounds 214, 218, 219
 C-glycoside rearrangement 179–80
 glycosyl acetates 160
 O-glycosylation 160, 177–8
 glycosyl fluoride 177–8
 glyoxylate-ene reaction 101–2, 104, 106, 107–8
 glyoxylates 99–100, 121, 122, 125, 230
 gold(I) reagents 169–73
 bisphosphine complex 169, 170, 171, 172–3
 Grignard reagents 6, 96, 97–9
- hafnocene complexes 177
 activation of carbon–fluorine bond 177–8
 activation of ethers 179–80
 carbonyl activation 180–1
 haloalkanes: activation 225–7
 α -haloketones 204
 Henry (nitroaldol) reaction 217, 218–19
 hexamethyldisiloxane 160
 hexanal 163–4
 HMG-Co A inhibitors 121
 holmium triflate 215
 homoallyl alcohols 53, 163, 192
 homoallyl ethers 163
 Hosomi–Sakurai reaction, *see*
 Sakurai–Hosomi reaction
 hydrogen cyanide 130
 hydrophosphonylation 217, 218
 hydrosilylation 35, 199
threo- β -hydroxy- α -amino acids 169, 171
 β -hydroxy- α -amino aldehydes/ketones 169
 β -hydroxy aldehydes 26
 2-(1-hydroxycyclohexyl)-1-phenyl-1-
 propanone 207–8
 β -hydroxy esters 52
 α -hydroxy esters 101, 102–3
 hydroxy-pyridines 78
- imines
 asymmetric reactions with chiral boron
 reagents 57–61
 hydrophosphonylation 217, 218
 imino groups 79
 iodomethane complexes 225–6
 iodotrimethylsilane 159, 160–2, 163–7
 (*R*)-(-)-ipsdienol 103, 104
- iron complexes
 ferrocenium ion/ferrocene 228, 229
 iron-bisoxazoline 230, 231, 232–3
 iron-bisphosphite 234, 235–6
 iron-bissulfoxide 230, 231
 isocarbacycline analogues 105
 isocyanides 169–73
 α -isocyano Weinreb amide 169
- Kaminsky polymerization 182
 ketene bis(trialkylsilyl) acetals 153, 154
 ketene silyl acetals (KSA)
 aldol and aldol-type reactions 52, 71, 118,
 119, 208–9, 210, 238–42
 with imines 33–4, 58, 59
 asymmetric protonation 153, 154
 Michael reactions 208, 209, 210, 213–14
 α -ketoesters 140–2
 ketones
 acetalization 160, 189
 cyanosilylation 215
 and diethylzinc 78
 ether formation 164–5
 hydrosilylation 35
 and KSA 209
 reduction 203–4
 see also in particular aldol and aldol-type
 reactions; alkylation; allylation; *named*
 ketones
 kinetic optical resolution 106–7
 KSA, *see* ketene silyl acetals
- β -lactone 73–4, 117
 lanthanide(III) reagents **203–23**
 alkoxides 220
 hydride complexes 220
 isopropoxides 219–20
 polymerization reactions 220–1
 lanthanum(III) reagents 217–19
 Lewis acid-assisted chiral Brønsted acid
 (LBA) 153–4
 Lewis acid-assisted high-speed living anionic
 polymerization 246–60
 Lewis bases 152
 ligand modification/redistribution 2, 94–6
 lithium alkynides 11
 lithium carbenoids 11
 lithium diallylcuprate 12
 lithium enolates 204
 lithium perchlorate 65–7, 71
 living polymerization 221, 246
 lupin alkaloids 199, 200
 lutetium compounds 215, 220
- MAD/MAT 6–10, 15–16
 magnesium bromide 65, 67–8, 71–3

- magnesium bromide etherate 73–4
 magnesium(II) compounds 65, 67–74
 chiral 68–70
 malate derivatives: synthesis of 2-substituted 140–3
 manganese(V) complexes 226, 238
 Meerwein–Ponndorff–Verley (MPV)
 reductions 219–20
 metallocenes 228–9
 metalloporphyrin complexes 246–7
 methacrolein
 and 1,3-dienol derivatives 122–4
 and cyclopentadiene 39, 40, 44, 45, 234–8
 4-methoxy-1-nonene 163–4
p-methoxyacetophenone 187–8
 1-methoxydienes 121, 122
N-methoxy-*N*-methyl- α -isocyanacetamide 169
 (*S*)-1-methyl-2-[(*N*-1-naphthylamino)methyl]-pyrrolidine 143
 methylacrylaldehyde 81–3
 methyl acrylate 45, 65, 66–7
 methylaluminium 5,10,15,20-tetraphenylporphine 249–54
 methylaluminium bis(2,6-diphenylphenoxide) 254, 255–6
 methylaluminium bis(2,6-di-*tert*-butyl-4-alkylphenoxide) (MAT and MAD) 6–10, 15–16
 methylaluminium bis(2,6-di-*tert*-butyl-4-bromophenoxide) (MABR) 22, 23–8
 methylaluminium bis(2,6-di-*tert*-butyl-4-methylphenolate) 258–60
 methylaluminium bis(2,6-di-*tert*-butyl-4-methylphenoxide) 249, 251–2
 methylaluminium bis(4-bromo-2,6-di-isopropylphenoxide) (MAIP) 28
 methylaluminium diphenoxides 247
 methylaluminium tetramethylporphyrin 248
 methylaluminoxane 182–4
 β -methyl cyclohexanone 227
 methyl isocyanoacetate 174–5
 methyl isopropylglyoxylate 140
 methylolithium 6–10
 methyl methacrylate: polymerization 220–1, 248, 249–50
 methyl phenylglyoxylate 140, 143
 methyl pyruvate 140, 142
 mevinolin 121
 Michael reactions 3, 65, 66, 130
 europium(III) complex catalysis 208, 209, 210, 213–14
 lanthanum(III) reagent catalysis 217
 ytterbium triflate catalysis 214
 molybdenum complexes 228, 229
 monoalkylzinc alkoxides 78
 monosaccharides 121
 Mukaiyama aldol reactions 2
 and boron reagents 32–3, 52–3, 56–7
 and copper(II) complexes 241–2
 and lithium perchlorate 65, 66, 71
 and rhodium complexes 238, 239
 and ruthenium complexes 238, 239–40
 and scandium(III) triflate 194–6
 and silicon(IV) reagents 161–3
 and titanium complexes 113–18
 see also silyl enol ethers

 naphthol derivatives: acylation 186
 naphthoquinone derivatives 122–3
 naproxen 154
 neodymium triflate 214
 nickel complexes 78, 85–6
 nitroaldol (Henry) reactions 217, 218–19
 nitroalkanes 217
p-nitrobenzaldehyde 209
 nitrones 69–70
 non-linear effects (NLE) 107–8, 125–6
 nucleoside synthesis 160, 161

 olefins 102, 104, 112, 220
 oligosaccharides 160, 161, 177
 onium salts 249
 optical resolution, kinetic 106–7
 organic synthesis, role of Lewis acid reagents in 1–4
 organoaluminium compounds, *see* aluminium reagents
 organocerium reagents, *see* cerium(III) reagents
 organolithiums 6–15
 organosilicon reagents, *see* silicon(IV) reagents
 organozincs, *see* zinc(II) reagents
 orthoesters 161
 osmium tetroxide 143
 oxaborolidines 55
 oxazolidinones 68–9, 119, 196–9, 232
 oxazolines 169
 oxetanes: polymerization 248, 258–60
 oxiranes 215, 248
 oxo(salen)manganese(V) complexes 238
 oxovanadium(IV) complexes 230

 palladium complexes 240
 (*S*)-1-pentyl-2-[(piperidin-1-yl)methyl]pyrrolidine 140, 142
 phenol derivatives: acylation 186
 1-phenyl-1-[(trimethylsilyl)oxy]ethylene 239–40
 (*S*)-1-phenyl-3-butene-1-ol 174–5
 3-phenylacetylacetone 214
 1-phenylethanol 227

- phenylsilane 199
 phosphine-coordinated complexes 169–75
 phosphines 229
 phosphite-iron(II) complexes 234, 235–6
 phthalocyanine-aluminium complexes 248
trans-piperylene 66–7
 poly(δ -valerolactone) 254–8
 polymerization 182, 220–1
 synthesis of polymers with controlled molecular weight **245–61**
 poly(methyl methacrylate) 249–54
 polyoxetane 248, 258–60
 praseodymium compounds 218
 prenyllithium 14
 prochiral enol derivatives 153–4
 prolinol derivatives 45
 (*S*)-1-propyl-2-[(piperidin-1-yl)methyl]pyrrolidine 145
i-propylcerium 206
 protonation, asymmetric 153–4
 pybox 236, 237–8, 240, 241–2
 pyridinecarbaldehydes 78
 pyrrolizidine alkaloids 140
- racemic catalysts, enantiomer-selective activation 108–11
 racemic substrates, resolution 106–7
 rhenium complexes 226–7
 rhodium-diphosphine complexes 238
 ruthenium complexes 225–6, 229, 238–40
- saccharides 121, 192
 Sakurai–Hosomi reaction
 BINOL-Ti catalysed 99–100
 CAB catalysed 53–5
 silicon(IV) mediated 163
 samarium compounds 219, 220
 scandium(III) triflate 185–99
 acetalization 189–90
 acetylation and esterification of alcohol 190–2
 allylation with tetra-allyltin 192–4
 Diels–Alder reaction 196–9
 Friedel–Crafts reaction 185–9
 Mukaiyama aldol reaction 194–6
 Schiff base-aluminium complex 248
 Schiff base-titanium complex 118, 130
 selenides 103
 Sharpless asymmetric epoxidation 26
 1,3-sigmatropic rearrangements 65
 silanes 53, 99–101, 159, 163
 silatropic ene pathway 115–17
 silicates 159
 silicon(IV) reagents **159–68**
 carbon–carbon bond formation 161–4
 carbon–hydrogen bond formation 164–7
 carbon–nitrogen bond formation 160–1
 carbon–oxygen bond formation 160–1
 β -siloxy aldehydes 26–7
 4-siloxy cyclopentenone 209, 210
 siloxy dienes 67
 silver hexafluoroarsenate 180, 181–2
 silver(I) reagents **169–76**
 BINAP-complex 173–5
 bis(phosphino)ferrocene complex 169, 170, 171
 silver perchlorate 177–8, 179, 180
 silyl cyanides 208
 silyl enolates 161–2
 silyl enol esters 153
 silyl enol ethers 214–16
 and α,β -enones 214
 asymmetric aldol reactions 138–40
 catalytic asymmetric reactions 147–51
 synthesis of 1,2-diol derivatives 143–7
 synthesis of 2-substituted malate derivatives 140–3
 asymmetric protonation 153–4
 see also Mukaiyama aldol reactions
 sodium borohydride–cerium chloride system 203–6
 sodium naphthalide 246
 stannanes 71–3, 99–100
 steroidal ketones 7, 8
 trans-stilbene oxide, rearrangement of 24–6
 styrene 246
- TADDOL 83–4
 TADDOL-titanium complexes
 addition of alkylzincs 78, 96–9
 cyanohydrin formation 129–30
 cyclopropanation of alcohol 86–7
 ene-cyclization 112
 and oxazolidinone derivatives of acrylates 119–21
 tetra-allyltin 192–4
 tetraethylammonium chloride 258
 (5,10,15,20-tetraphenylporphinato)aluminium methoxide 254–8
 β -thioalcohols 78
 thioanisole 186
 thulium triflate 215
 tin(II) alkoxide 148
 tin(II) monoalkoxymonotriflate 152, 153
 tin(II) oxide 150
 tin(II) triflate 138–9, 140–51
 tin Lewis acids **137–57**
 asymmetric aldol reactions 138–51
 asymmetric cyanation reaction 151–3
 asymmetric protonation 153–4
 preparation of chiral tin(II) 138
 properties 137–8
 tin tetrachloride 153

- Tishchenko reaction 220
 titanium complexes, chiral **93–136**
 aldol reactions 113–18
 asymmetric catalytic desymmetrization 105–6
 carbonyl addition reactions 96–101
 carbonyl-ene reaction 101–5
 cyanohydrin formation 129–30
 enantioselective activation of racemic catalysts 108–11
 ene cyclization 111–13
 hetero-Diels–Alder reaction 118–29
 kinetic optical resolution 106–7
 positive non-linear effect of non-racemic catalysts 107–8
 preparative procedures 94–5
 titanium-disulfonamido system 78, 79, 86–7, 97
 titanium-TADDOL, *see* TADDOL-titanium complexes
 titanocene dichloride 229
 transhydrocyanation 220
 transition-metal catalysts 225–43
 activation of haloalkanes and carbonyl compounds 225–7
 asymmetric aldol condensations 238–42
 asymmetric Diels–Alder reactions
 achiral catalysts 228–9
 chiral catalysts 229–38
 see also named catalysts/reagents
 trialkylaluminium compounds 21, 247
 trialkyl amines 78
 triarylborons 248
 tributylaluminium 21
 tributyltin fluoride 138, 140, 141, 142, 143
 trifluoroacetic anhydride 55
 3,4,5-trifluorophenylboronic acid 31
 as amidation catalyst 37–9
 tri-isobutylaluminium 248
 trimethylaluminium 6, 17
 cis-1,2,3-trimethylpiperidine 196
 trimethylsilane 164–7
 1-trimethylsiloxy-1-cyclohexene 195–6
 1-trimethylsiloxy-1-phenylethene 162–3
 trimethylsilyl cyanide 129–30, 152
 trimethylsilyl enol ether 148, 153
 trimethylsilylpropynal 83–4
 trimethylsilyl triflate 148, 149, 150
 trimethylsilyl trifluoromethanesulfonate 159, 160–5
 triphenylaluminium 248
 triphenylboron 248
 triphenylphosphine 226
 tris(pentafluorophenyl)boron 31, 32–5, 248
 tungsten nitrosyl complex 228
 δ -valerolactone: polymerization 221, 254–8

 vanadocenium complex 229
 verrucaric acid 99
 vinylic ethers 21–3, 215
 vinylic sulfides 103
 ytterbium(III) reagents 214–16
 yttrium(III) compounds 199–200

 zinc aminoalkoxide 78
 zinc(II) reagents **77–91**
 zinc iodide 152
 zirconocene complexes 177
 carbonyl activation 180–2
 epoxide rearrangement 179
 reductive coupling of allenes and alkynes 182–4
 zirconocene dichloride 229