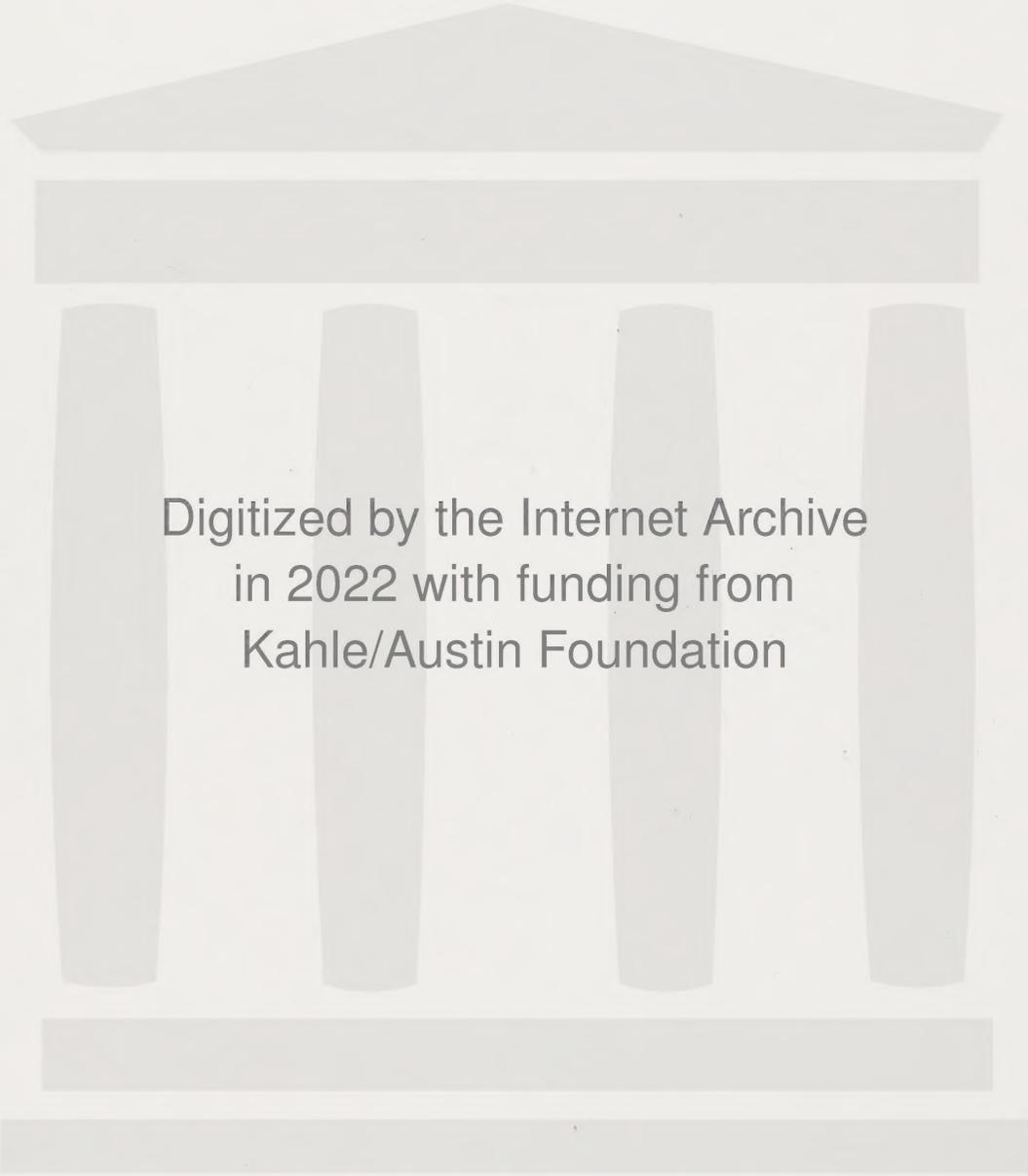


Wilfred L. F. Armarego • Christina Li Lin Chai

# PURIFICATION OF LABORATORY CHEMICALS

SEVENTH EDITION





Digitized by the Internet Archive  
in 2022 with funding from  
Kahle/Austin Foundation

<https://archive.org/details/purificationofla0000arma>

***PURIFICATION  
OF  
LABORATORY CHEMICALS***

***Seventh Edition***

## DEDICATION

The past years have seen the loss of two staunch champions for the both of us — the late Athel L.J. Beckwith FAA, FRS from the Research School of Chemistry, ANU (scientist extraordinaire, mentor and friend) as well as Mrs Joan E. Smith, Research School of Chemistry ANU (a dedicated and amazing librarian, and a strong supporter of this book and its previous editions. In her eyes, we could do no wrong). The world is a little less bright in their absence. We dedicate this book to their memories.

W.L.F. Armarego & C.L.L. Chai

# Purification of Laboratory Chemicals

Seventh Edition

**Wilfred L. F. Armarego**

The John Curtin School of Medical Research  
ANU College of Medicine, Biology, and Environment  
The Australian National University, Canberra, A.C.T.  
AUSTRALIA

**Christina Li Lin Chai**

Department of Pharmacy, National University of Singapore (NUS)  
and  
Institute of Chemical and Engineering Sciences  
Agency for Science, Technology, and Research (A\*STAR)  
SINGAPORE



ELSEVIER

AMSTERDAM • BOSTON • HEIDELBERG • LONDON  
NEW YORK • OXFORD • PARIS • SAN DIEGO  
SAN FRANCISCO • SINGAPORE • SYDNEY • TOKYO



MAIN  
TP  
156  
P83  
P47  
2013

Butterworth-Heinemann is an imprint of Elsevier  
The Boulevard, Langford Lane, Kidlington, Oxford OX5 1GB, UK  
225 Wyman Street, Waltham, MA 02451, USA

Sixth edition 2008  
Seventh edition 2013

Copyright © 2013 Elsevier Inc. All rights reserved

No part of this publication may be reproduced, stored in a retrieval system or transmitted in any form or by any means electronic, mechanical, photocopying, recording or otherwise without the prior written permission of the publisher

Permissions may be sought directly from Elsevier's Science & Technology Rights Department in Oxford, UK: phone (+44) (0) 1865 843830; fax (+44) (0) 1865 853333; email: [permissions@elsevier.com](mailto:permissions@elsevier.com). Alternatively you can submit your request online by visiting the Elsevier web site at <http://elsevier.com/locate/permissions>, and selecting *Obtaining permission to use Elsevier material*

#### Notice

No responsibility is assumed by the publisher for any injury and/or damage to persons or property as a matter of products liability, negligence or otherwise, or from any use or operation of any methods, products, instructions or ideas contained in the material herein. Because of rapid advances in the medical sciences, in particular, independent verification of diagnoses and drug dosages should be made

#### British Library Cataloguing in Publication Data

A catalogue record for this book is available from the British Library

#### Library of Congress Cataloging-in-Publication Data

A catalog record for this book is available from the Library of Congress

ISBN: 978-0-12-382161-4

For information on all Butterworth-Heinemann publications visit our web site at [books.elsevier.com](http://books.elsevier.com)

Printed and bound in the US

12 13 14 15 16 10 9 8 7 6 5 4 3 2 1

Working together to grow  
libraries in developing countries

[www.elsevier.com](http://www.elsevier.com) | [www.bookaid.org](http://www.bookaid.org) | [www.sabre.org](http://www.sabre.org)

ELSEVIER BOOK AID International Sabre Foundation

Cooper Union Library

FFR 18 2014

## Preface to the Seventh Edition

THE SALES of the sixth edition, which appeared in April 2009, were high by about October 2009, and one of us (WLFA) was approached by Ms Melanie Benson, Editorial Project Manager of Elsevier Science & Technology Books (who was mainly responsible for the production of the 6th edition), about writing a 7th edition within 2–3 years. In the past, 6–9 years were allowed to lapse between editions. However, the attraction this time, was that we were allowed to increase the size of the work by up to 249 pages. This has given us the opportunity to update all the previous chapters in the light of current thinking on safety (personal and environmental), and to introduce two new chapters. The award of five Nobel Prizes in the past ten years or so in Chemistry and one in Physics, of which three were awarded for work on *Catalysis and the catalytic process* (2001, 2005 and 2010), and two for work on *Nanomaterials and Nanotechnology* (1996, 2010) have prompted us to write a new chapter on *Catalysts* and a new chapter on *Nanomaterials and Nanotechnology*. Chemical suppliers have now made commercially available a large number of catalysts as well as many nanomaterials of various sorts. Since the number of commercially available catalysts are currently considerably larger than that of nanomaterials, the chapter on Catalysts is larger than that on Nanomaterials and Nanotechnology, and had to be divided into two parts. The availability, preparation and purification of a large range of these, are presented in these chapters. The other chapters have been updated and expanded, also in keeping with the purpose of all previous editions which is to provide information for the purification of commercially available laboratory materials. Of course, the General Subjects Index and the Chemical Abstracts Registry Numbers (CASRNs) Index increased in size accordingly. Much of the cross-referencing is done *via* CASRNs and a page of how to use this book through these is included before Chapter 1 to assist the reader, not only to locate the pages where the required CASRNs are to be found, but also to let the reader know whether a particular substance is included in this work. CASRNs can be readily obtained from chemical catalogues or from SciFinder.

We would like to acknowledge Professor Martin Banwell FAA (Director, Research School of Chemistry, ANU) for his generosity in allowing the use of IT services, and to Dr Emil Mittag (Research School of Chemistry, ANU) for editing and updating the *Macro* program used for making the CASRNs Index.

We are greatly indebted to many people who have made valuable and constructive comments, and indicated errors, in previous editions. The continued help from Joe Papa of BS MS (EXAXOL in Clearwater, Florida, USA) with the preparation and purification of several inorganic compounds, particularly beryllium and cerium, is appreciated and gratefully acknowledged.

One of us (W.L.F.A) owes a debt of gratitude to Professor Jill E. Gready (John Curtin School of Medical Research, ANU) for her continued encouragement and for strongly supporting a Visiting Fellowship over a period of many years.

We thank Dr Pauline M. Armarego for assistance in the painstaking task of proofreading and correcting typographical errors as well as checking the General Index and the CASRNs Index.

We thank the ANU library and its staff, both on campus and at the ANU Print Repository, for their prompt and unfailing assistance.

W.L.F. Armarego & C.L.L. Chai

May 2012

### Preface to the First Edition

WE BELIEVE that a need exists for a book to help the chemist or biochemist who wishes to purify the reagents she or he uses. This need is emphasised by the previous lack of any satisfactory central source of references dealing with individual substances. Such a lack must undoubtedly have been a great deterrent to many busy research workers who have been left to decide whether to purify at all, to improvise possible methods, or to take a chance on finding, somewhere in the chemical literature, methods used by some previous investigators.

Although commercially available laboratory chemicals are usually satisfactory, as supplied, for most purposes in scientific and technological work, it is also true that for many applications further purification is essential.

With this thought in mind, the present volume sets out, first, to tabulate methods, taken from the literature, for purifying some thousands of individual commercially available chemicals. To help in applying this information, two chapters describe the more common processes currently used for purification in chemical laboratories and give fuller details of new methods which appear likely to find increasing application for the same purpose. Finally, for dealing with substances not separately listed, a chapter is included setting out the usual methods for purifying specific classes of compounds.

To keep this book to a convenient size, and bearing in mind that its most likely users will be laboratory-trained, we have omitted manipulative details with which they can be assumed to be familiar, and also detailed theoretical discussion. Both are readily available elsewhere, for example in Vogel's very useful book **Practical Organic Chemistry** (Longmans, London, 3rd ed., 1956), or Fieser's **Experiments in Organic Chemistry** (Heath, Boston, 3rd ed., 1957).

For the same reason, only limited mention is made of the kinds of impurities likely to be present, and of the tests for detecting them. In many cases, this information can be obtained readily from existing monographs.

By its nature, the present treatment is not exhaustive, nor do we claim that any of the methods taken from the literature are the best possible. Nevertheless, we feel that the information contained in this book is likely to be helpful to a wide range of laboratory workers, including physical and inorganic chemists, research students, biochemists, and biologists. We hope that it will also be of use, although perhaps to only a limited extent, to experienced organic chemists.

We are grateful to Professor A. Albert and Dr D.J. Brown for helpful comments on the manuscript.

D.D.P., W.L.F.A. & D.R.P.  
1966

### Preface to the Second Edition

SINCE the publication of the first edition of this book, there have been major advances in purification procedures. Sensitive methods have been developed for the detection and elimination of progressively lower levels of impurities. Increasingly stringent requirements for reagent purity have gone hand-in-hand with developments in semiconductor technology, in the preparation of special alloys and in the isolation of highly biologically active substances. The need to eliminate trace impurities at the micro- and nanogram levels has placed greater emphasis on ultrapurification technique. To meet these demands the range of purities of laboratory chemicals has become correspondingly extended. Purification of individual chemicals thus depends more and more critically on the answers to two questions—Purification from what, and to what permissible level of contamination. Where these questions can be specifically answered, suitable methods of purification can usually be devised.

Several periodicals devoted to ultrapurification and separations have been started. These include "Progress in Separation and Purification" (vol. 1) Ed. E.S. Perry, Wiley-Interscience, New York, vols. 1-4, 1968-1971, and **Separation and Purification Methods**, Ed. E.S. Perry and C.J. van Oss, Marcel Dekker, New York, vol. 1, 1973. Nevertheless, there still remains a broad area in which a general improvement in the level of purity of many compounds can be achieved by applying more or less conventional procedures. The need for a convenient source of information on methods of purifying available laboratory chemicals was indicated by the continuing demand for copies of this book even though it had been out of print for several years.

We have sought to revise and update this volume, deleting sections that have become more familiar or less important, and incorporating more topical material. The number of compounds in Chapters 3 and 4 have been increased appreciably. Also, further details in purification and physical constants are given for many compounds that were listed in the first edition.

We take this opportunity to thank users of the first edition who pointed out errors and omissions, or otherwise suggested improvements or additional material that should be included. We are indebted to Mrs S. Schenk who emerged from retirement to type this manuscript.

D.D.P., W.L.F.A. & D.R.P.  
1980

### Preface to the Third Edition

THE CONTINUING demand for this monograph and the publisher's request that we prepare a new edition are an indication that **Purification of Laboratory Chemicals** fills a gap in many chemists' reference libraries and laboratory shelves. The present volume is an updated edition that contains significantly more detail than the previous editions, as well as an increase in the number of individual entries and a new chapter.

Additions have been made to Chapters 1 and 2 in order to include more recent developments in techniques (e.g. Schlenk-type, *cf* p. 10), and chromatographic methods and materials. Chapter 3 still remains the core of the book, and lists in alphabetical order relevant information on *ca* 4000 organic compounds. Chapter 4 gives a smaller listing of *ca* 750 inorganic and metal-organic substances, and makes a total increase of *ca* 13% of individual entries in these two chapters. Some additions have also been made to Chapter 5.

We are currently witnessing a major development in the use of physical methods for purifying large molecules and macromolecules, especially of biological origin. Considerable developments in molecular biology are apparent in techniques for the isolation and purification of key biochemicals and substances of high molecular weight. In many cases something approaching homogeneity has been achieved, as evidenced by electrophoresis, immunological and other independent criteria. We have consequently included a new section, Chapter 6, where we list upwards of 100 biological substances to illustrate their current methods of purification. In this chapter the details have been kept to a minimum, but the relevant references have been included.

The lists of individual entries in Chapters 3 and 4 range in length from single-line entries to *ca* one page or more for solvents such as acetonitrile, benzene, ethanol and methanol. Some entries include information such as likely contaminants and storage conditions. More data referring to physical properties have been inserted for most entries [i.e. melting and boiling points, refractive indexes, densities, specific optical rotations (where applicable) and UV absorption data]. Inclusion of molecular weights should be useful when deciding on the quantities of reagents needed to carry out relevant synthetic reactions, or preparing analytical solutions. The Chemical Abstracts registry numbers have also been inserted for almost all entries and should assist in the precise identification of the substances.

In the past ten years laboratory workers have become increasingly conscious of safety in the laboratory environment. We have therefore in three places in Chapter 1 (pp. 3 and 33, and bibliography p. 52) stressed more strongly the importance of safety in the laboratory. Also, where possible, in Chapters 3 and 4 we draw attention to the dangers involved with the manipulation of some hazardous substances.

The worldwide facilities for retrieving chemical information provided by the Chemical Abstract Service (CAS on-line) have made it a relatively easy matter to obtain CAS registry numbers of substances, and most of the numbers in this monograph were obtained *via* CAS on-line. We should point out that two other available useful files are CSChem and CSCORP, which provide, respectively, information on chemicals (and chemical products) and addresses and telephone numbers of the main branch offices of chemical suppliers.

The present edition has been produced on an IBM PC and a Laser Jet printer using the **Microsoft Word (4.0)** word-processing program with a set stylesheets. This has allowed the use of a variety of fonts and font sizes which has made the presentation more attractive than in the previous edition. Also, by altering the format and increasing slightly the sizes of the pages, the length of the monograph has been reduced from 568 to 391 pages. The reduction in the number of pages has been achieved in spite of the increase of *ca* 15% of total text.

We extend our gratitude to the readers whose suggestions have helped to improve the monograph, and to those who have told us of their experiences with some of the purifications stated in the previous editions, and in particular with the hazards that they have encountered. We are deeply indebted to Dr M.D. Fenn for the several hours that he has spent on the terminal to provide us with a large number of CAS registry numbers.

This monograph could not have been produced without the expert assistance of Mr David Clarke who has spent many hours loading the necessary fonts in the computer, and for advising one of the authors (W.L.F.A.) on how to use them together with the idiosyncrasies of Microsoft Word.

D.D.P. & W.L.F.A.  
1988

#### Preface to the Fourth Edition

THE AIMS of the first three editions, to provide purification procedures of commercially available chemicals and biochemicals from published literature data, are continued in this fourth edition. Since the third edition in 1988 the number of new chemicals and biochemicals that have been added to most chemical and biochemical catalogues have increased enormously. Accordingly there is a need to increase the number of entries with more recent useful reagents and chemical and biochemical intermediates. With this in mind, together with the need to reorganise and update general purification procedures, particularly in the area of biological macromolecules, as well as the time lapse since the previous publication, this fourth edition of **Purification of Laboratory Chemicals** has been produced. Chapter 1 has been reorganised with some updating, and by using a smaller font it was kept to a reasonable number of pages. Chapters 2 and 5 were similarly altered and have been combined into one chapter. Eight hundred and three hundred and fifty entries have been added to Chapters 3 (25% increase) and 4 (44% increase), respectively, and four hundred entries (310% increase) were added to Chapter 5 (Chapter 6 in the Third Edition), making a total of 5700 entries—all resulting in an increase from 391 to 529 pages, i.e., by *ca* 35%.

Many references to the original literature have been included remembering that some of the best references happened to be in the older literature. Every effort has been made to provide the best references, but this may not have been achieved in all cases. Standard abbreviations, listed on page 1, have been used throughout this edition to optimise space, except where no space advantage was achieved, in which cases the complete words have been written down to improve the flow of the sentences.

With the increasing facilities for information exchange, chemical, biochemical and equipment suppliers are making their catalogue information available on the Internet; e.g., Aldrich-Fluka-Sigma catalogue information is available on the World Wide Web by using the address <http://www.sigma.sial.com> and GIBCO BRL catalogue information from <http://www.lifetech.com> as well as on CD-ROMS which are regularly

updated. Facility for enquiring about, ordering and paying for items is available *via* the Internet. CAS on-line can be accessed on the Internet, and CAS data is available now on CD-ROM. Also biosafety bill boards can similarly be obtained by sending SUBSCRIBE SAFETY John Doe at the address "listserv@uvmvm.uvm.edu", SUBSCRIBE BIOSAFETY at the address "listserv@mitvma.mit.edu", and SUBSCRIBE RADSFAF at the address "listserv@romulus.ehs.uiuc.edu"; and the Occupational, Health and Safety information (Australia) is available at the address "http://www.safework.gov.au". Sigma-Aldrich provided Material Safety data sheets on CD-ROMs.

It is with much sadness that Dr Douglas D. Perrin was unable to participate in the preparation of the present edition due to illness. His contributions towards the previous editions have been substantial, and his drive and tenacity have been greatly missed.

The Third Edition was prepared on an IBM-PC, and the previous IBM files were converted into Macintosh files. These have now been reformatted on a Macintosh LC575 computer, and all further data to complete the Fourth Edition were added to these files. The text was printed with a Hewlett-Packard 4MV -600dpi Laser Jet printer, which gives a clearer resolution.

I thank my wife Dr Pauline M. Armarego, also an organic chemist, for the arduous and painstaking task of entering the new data into the respective files, and for the numerous hours of proofreading as well as the corrections of typographic errors in the files. I should be grateful to my readers for any comments, suggestions, amendments and criticisms which could, perhaps, be inserted in the second printing of this edition.

W.L.F. Armarego, 30 June 1996

### Preface to the Fifth Edition

THE DEMAND for **Purification of Laboratory Chemicals** has not abated since the publication of the fourth edition as evidenced by the number of printings and the sales. The request by the Editor for a fifth edition offered an opportunity to increase the usefulness of this book for laboratory purposes. It is with deep regret that mention should be made that Dr Douglas D. Perrin had passed away soon after the fourth edition was published. His input in the first three editions was considerable, and his presence has been greatly missed. A fresh, new and young outlook was required in order to increase the utility of this book, and it is with great pleasure that Dr Christina L.L. Chai, a Reader in Chemistry and leader of a research group in organic and bio-organic chemistry, has agreed to coauthor this edition. The new features of the fifth edition have been detailed below.

Chapters 1 and 2 have been reorganised and updated in line with recent developments. A new chapter on the Future of Purification has been added. It outlines developments in syntheses on solid supports, combinatorial chemistry as well as the use of ionic liquids for chemical reactions and reactions in fluorinated media. These technologies are becoming increasingly useful and popular, so much so that many future commercially available substances will most probably be prepared using these procedures. Consequently, knowledge of their basic principles will be helpful in many purification methods of the future.

Chapters 4, 5 and 6 (3, 4 and 5 in the 4th ed.) form the bulk of the book. The number of entries has been increased to include the purification of many recent commercially available reagents that have become more and more popular in the syntheses of organic, inorganic and bio-organic compounds. Several purification procedures for commonly used liquids, e.g., solvents, had been entered with excessive thoroughness, but in many cases the laboratory worker only requires a simple, rapid but effective purification procedure for immediate use. In such cases a **rapid purification** procedure has been inserted at the end of the respective entry, and should be satisfactory for most purposes. With the increased use of solid phase synthesis, even for small molecules, and the use of reagents on solid support (e.g., on polystyrene) for reactions in liquid media, compounds on solid support have become increasingly commercially available. These have been inserted at the end of the respective entry and have been listed in the General Index together with the above rapid purification entries.

A large number of substances are ionisable in aqueous solutions, and knowledge of their ionisation constants, stated as pK (pKa) values, can be of importance not only in their purification but also in their reactivity. Literature values of the pK's have been inserted for ionisable substances, and where values could not be found they were estimated (pK<sub>Est</sub>). The estimates are usually so close to the true values as not to affect the purification process or the reactivity seriously. The book will thus be a good compilation of pK values for ionisable substances.

Almost all the entries in Chapters 4, 5 and 6 have CAS (Chemical Abstract Service) Registry Numbers to identify them, and these have been entered for each substance. Unlike chemical names which may have more than one synonymous name, there is only one CAS Registry Number for each substance (with only a few exceptions, e.g., where a substance may have another number before purification, or before determination of absolute configuration). To simplify the method for locating the purification of a substance, a CAS Registry Number Index with the respective page numbers has been included after the General Index at the end of the book. This will also provide the reader with a rapid way to see if the purification of a particular substance has been reported in the book. The brief General Index includes page references to procedures and equipment, page references to abbreviations of compounds, e.g., TRIS, as well as the names of substances for which a Registry Number was not found.

Website references for distributors of substances or/and of equipment have been included in the text. However, since these may be changed in the future we must rely on the suppliers to inform users of their change in website references.

We wish to thank readers who have provided advice, constructive criticism and new information for inclusion in this book. We should be grateful to our readers for any further comments, suggestions, amendments and criticisms which could, perhaps, be inserted in a second printing of this edition. In particular, we thank Professor Ken-chi Sugiura (Graduate School of Science, Tokyo Metropolitan University, Japan) who has

provided us with information on the purification of several organic compounds from his own experiences, and Joe Papa BS MS (EXAXOL in Clearwater, Florida, USA) who has provided us not only with his experiences in the purification of many inorganic substances in this book, but also gave us his analytical results on the amounts of other metal impurities at various stages of purification of several salts. We thank them graciously for permission to include their reports in this work. We express our gratitude to Dr William B. Cowden for his generous advice on computer hardware and software over many years and for providing an Apple LaserWriter (16/600PS) which we used to produce the master copy of this book. We also extend our sincere thanks to Dr Bart Eschler for advice on computer hardware and software and for assistance in setting up the computers (iMac and eMac) used to produce this book.

We thank Dr Pauline M. Armarego for assistance in the painstaking task of entering data into respective files, for many hours of proofreading, correcting typographical errors and checking CAS Registry Numbers against their respective entries.

One of us (W.L.F.A) owes a debt of gratitude to Dr Desmond (Des) J. Brown of the Research School of Chemistry, ANU, for unfailing support and advice over several decades and for providing data that was difficult to acquire not only for this edition but also for the previous four editions of this book.

One of us (C.L.L.C) would especially like to thank her many research students (past and present) for their unwavering support, friendship and loyalty, which enabled her to achieve what she now has. She wishes also to thank her family for their love, and would particularly like to dedicate her contribution towards this book to the memory of her brother Andrew who had said that he should have been a scientist.

We thank Mrs. Joan Smith, librarian of the Research School of Chemistry, ANU, for her generous help in many library matters, which has made the tedious task of checking references more enduring.

W.L.F. Armarego & C.L.L. Chai  
November 2002

#### Preface to the Sixth Edition

THERE IS a continuing demand for the **Purification of Laboratory Chemicals** book, to the extent that the 5th edition which was published in early 2003 was carefully translated into Chinese (ISBN 978-7-5025-94367) by Ying-Jie Lin, Wei Liu, Hui-Ping Wang, Xiao-Bo Sun, Qing-Shan Li and Jun-Gang Cao from Jilin University (People's Republic of China) in 2007. In response to the demand, it was timely to update the 5<sup>th</sup> edition to include the more recently developed purification procedures, as well as add to the list of compounds for purification. The latter comprise some commercially available compounds that have gained usefulness and popularity in the past few years.

The first two chapters have been updated, sections of current interest have been expanded and new sections added. Chapter 3 has been rewritten so that areas of work that have lost popularity have been reduced in size or deleted and sections on recent, and now commonly adopted, technologies have been inserted. Chapters 4, 5 and 6 are now completely reorganized, and each is subdivided into several sections which will make it easier for the reader to locate compounds of similar classification. Chapter 4 is subdivided into aliphatic, alicyclic, aromatic and heterocyclic compounds, Chapter 5 has been subdivided into inorganic and metal-organic compounds, and Chapter 6 has been subdivided into amino acids and peptides, proteins, enzymes, DNA and RNA, carotenoids, carbohydrates, steroids and a miscellaneous section which includes small biologically active substances such as antibiotics, coenzymes, co-factors, lipids, phospholipids, polynucleotides and vitamins. Some useful compounds that have been added recently to commercial catalogues have been included in these three chapters. A large number of derivatives of previous entries with their physical properties and purifications have been inserted together with extensive referencing to the original literature including *Beilstein* references. This resulted in an increase in size of the 5th edition, in text and number of compounds, by over 20%. The purifications of some 7400 substances are described. As in the 5th edition, substance entries are in alphabetical order within subsections and each substance is defined by its Chemical Abstracts Service (CAS) Registry Number. An index of these numbers with their respective page numbers at the end of the book will make it possible to locate the purification of a desired substance readily and to check if the substance is contained in the book. For this purpose we thank Rodney Armarego for setting up a *Macro* on the MacBook Pro computer used for collating the CAS Registry Numbers for the index. There is also a General Index of Contents.

Website references of distributors of substances and/or of equipment have been included in the text. However, since these may change in the future, users should check for current websites of suppliers. The bibliographies have been updated, and websites of a few publishers and book suppliers have been included. Several texts with publication dates older than fifteen years have been deleted except for a few very useful textbooks which are out of print and where recent editions have not been produced. In these cases it is usually possible to obtain used copies from good suppliers of old books, for which there are several websites, e.g. visit Google under "old books suppliers"; also visit websites such as <<http://www.abebooks.com>>, <<http://www.betterworld.com/usedbooks>>, <<http://www.booksandcollectibles.com.au/index>>, <<http://www.ebay.com.au>>. Further information for almost every entry in Chapters 4, 5 and 6 of the 6th edition can be obtained from the references to the original literature, which are cited under each entry together with their respective *Beilstein* reference(s).

We thank readers who have provided advice, constructive criticism and new information. We are grateful for any further comments, suggestions, amendments and criticisms which could, perhaps, be inserted in a second printing of this edition. We thank Joe Papa BS MS (EXAXOL in Clearwater, Florida, USA) in particular for sharing his experiences on the purification of several inorganic substances in this and previous editions, and also for allowing us to use his analytical results on the amounts of metal impurities at various stages of purification of several salts.

We thank Dr Pauline M. Armarego for assistance in the painstaking task of entering data into respective files, for many hours of proofreading, correcting typographical errors and checking CAS Registry Numbers against their respective entries.

One of us (W.L.F.A) owes a debt of gratitude to Dr Desmond (Des) J. Brown of the Research School of Chemistry, ANU, for unfailing support and advice over several decades and for providing data that was difficult to acquire not only for this edition but also for the previous five editions of this book.

One of us (C.L.L.C) would like to acknowledge the support and friendship of her many research staff and students (past and present at ANU and A\*STAR). She especially thanks Drs Paul Huleatt, Paul Bernardo, Felicity Moore and Brendan Burkett for their unfailing faith in her, through chemical and personal journeys both in Singapore and Australia. The legacy of this book is for Kimberley and Victoria Tse because it is cool to be a scientist!

We thank Mrs Joan Smith, librarian of the Research School of Chemistry, ANU, for her generous help in many library matters which made the tedious task of checking references more enduring.

W.L.F. Armarego & C.L.L. Chai

November 2008

# CONTENTS

Preface to the Seventh Edition .....	v
Preface to the First Edition .....	vi
Preface to the Second Edition .....	vi
Preface to the Third Edition .....	vi
Preface to the Fourth Edition.....	vii
Preface to the Fifth Edition .....	viii
Preface to the Sixth Edition.....	ix

## CHAPTER 1

### COMMON PHYSICAL TECHNIQUES USED IN PURIFICATION

INTRODUCTION .....	1
THE QUESTION OF PURITY .....	1
SOURCES OF IMPURITIES .....	3
PRACTICES TO AVOID IMPURITIES .....	3
Cleaning practice .....	3
Silylation of glassware and plasticware .....	4
SAFETY PRECAUTIONS ASSOCIATED WITH THE PURIFICATION OF LABORATORY CHEMICALS .....	4
SOME HAZARDS OF CHEMICAL MANIPULATION IN PURIFICATION AND RECOVERY OF RESIDUES .....	5
Perchlorates and perchloric acid.....	5
Peroxides .....	6
Heavy-metal-containing explosives .....	6
Strong acids .....	6

Picric acid.....	6
Reactive halides and anhydrides .....	6
Salts and organic esters of some inorganic acids .....	6
Solvents .....	6
 SAFETY DISCLAIMER .....	 7
 METHODS OF PURIFICATION OF REAGENTS AND SOLVENTS.....	 7
 SOLVENT EXTRACTION AND DISTRIBUTION .....	 7
 DISTILLATION .....	 8
Techniques .....	9
Distillation of liquid mixtures .....	9
Types of distillation.....	10
<i>The distilling flask</i> .....	10
<i>Types of columns and packings</i> .....	10
<i>Types of condensers</i> .....	11
<i>Vacuum distillation</i> .....	12
<i>Spinning-band distillation</i> .....	12
<i>Steam distillation</i> .....	13
<i>Azeotropic distillation</i> .....	13
<i>Kügelrohr distillation</i> .....	14
<i>Isopiestic or isothermal distillation</i> .....	14
 RECRYSTALLISATION.....	 14
Techniques .....	14
<i>Filtration</i> .....	15
Solvents.....	16
<i>Choice of solvents</i> .....	16
<i>Water</i> .....	16
<i>Petroleum ethers</i> .....	17
<i>Mixed solvents</i> .....	17
Recrystallisation from the melt .....	17
Zone refining.....	18
 SUBLIMATION.....	 18
 CHROMATOGRAPHY .....	 18
Vapour phase chromatography (GC or gas-liquid chromatography).....	19
Liquid chromatography .....	19
<i>Adsorption chromatography</i> .....	19
<i>Thin layer chromatography (TLC)</i> .....	20
<i>Paper chromatography</i> .....	21

<i>Column chromatography</i> .....	21
<i>Graded adsorbents and solvents</i> .....	21
<i>Preparation and standardisation of alumina</i> .....	21
<i>Preparation of other adsorbents</i> .....	22
<i>Silica</i> .....	22
<i>Diatomaceous earth</i> .....	22
<i>Charcoal</i> .....	22
<i>Cellulose</i> .....	22
<i>Supelco products</i> .....	23
Flash Chromatography (FC and HPFC) .....	23
Paired-ion Chromatography (PIC) .....	23
Ion-exchange Chromatography .....	24
<i>Ion-exchange Resins</i> .....	24
<i>Ion-exchange celluloses and sephadex</i> .....	25
<i>Sephadex</i> .....	26
<i>Sepharose</i> .....	26
<i>Cellex CM</i> .....	26
<i>Cellex D</i> .....	26
<i>Crystalline hydroxylapatite</i> .....	26
Gel filtration .....	26
High performance liquid chromatography (HPLC) .....	27
Detectors .....	27
Columns for HPLC .....	27
Other types of liquid chromatography .....	28
Monolithic chromatography .....	28
Ultra performance liquid chromatography (UPLC) .....	28
Automated column chromatography .....	28
 ELECTROPHORESIS .....	 29
 DRYING .....	 29
Removal of solvents .....	29
Removal of water .....	29
Intensity and capacity of common desiccants .....	30
Suitability of individual desiccants .....	30
Molecular sieves .....	32
 PROPERTIES USEFUL IN PURIFICATION .....	 33
Spectroscopic .....	33
Ionisation constants — pK .....	34
pK and temperature .....	35
pK and solvent .....	36
Solubilities of gases in liquids .....	36

MISCELLANEOUS TECHNIQUES .....	37
Freeze-pump-thaw and purging .....	37
Vacuum lines, Schlenk and glovebox techniques .....	37
CHEMICAL AND BIOCHEMICAL SOURCES .....	37
USEFUL WEBSITES.....	38
TABLES .....	39
TABLE 1. SOME COMMON IMMISCIBLE OR SLIGHTLY MISCIBLE PAIRS OF SOLVENTS AT AMBIENT TEMPERATURES .....	39
FIGURE 1. NOMOGRAM.....	40
TABLE 2A. PREDICTED EFFECT OF PRESSURE ON BOILING POINT .....	41
TABLE 2B. PREDICTED EFFECT OF PRESSURE ON BOILING POINT .....	42
TABLE 3. HEATING BATHS .....	43
TABLE 4. WHATMAN FILTER PAPERS.....	43
TABLE 5. MICRO FILTERS .....	44
TABLE 6. COMMON SOLVENTS USED IN RECRYSTALLISATION .....	45
TABLE 7. PAIRS OF MISCIBLE SOLVENTS.....	45
TABLE 8. MATERIALS FOR COOLING BATHS .....	46
TABLE 9. LIQUIDS FOR STATIONARY PHASES IN GAS CHROMATOGRAPHY .....	47
TABLE 10. METHODS OF VISUALISATION OF TLC SPOTS .....	47
TABLE 11. GRADED ADSORBENTS AND SOLVENTS FOR CHROMATOGRAPHY.....	48
TABLE 12. REPRESENTATIVE ION-EXCHANGE RESINS USED IN CHROMATOGRAPHY .....	48
TABLE 13. MODIFIED FIBROUS CELLULOSES FOR ION-EXCHANGE CHROMATOGRAPHY .....	48
TABLE 14. BEAD FORM ION-EXCHANGE PACKAGINGS FOR CHROMATOGRAPHY .....	49
TABLE 15. SELECTED CHIRAL COLUMNS FOR CHROMATOGRAPHY ...	49
TABLE 16. LIQUIDS FOR DRYING PISTOLS .....	50
TABLE 17. VAPOUR PRESSURES (mm Hg) OF SATURATED AQUEOUS SOLUTIONS IN EQUILIBRIUM WITH SOLID SALTS .....	51
TABLE 18. DRYING AGENTS FOR CLASSES OF COMPOUNDS.....	52
TABLE 19. STATIC DRYING FOR SELECTED LIQUIDS (25°C).....	52
TABLE 20. FLUOROCHROMES.....	53
TABLE 21. RESIDUAL SOLVENT SIGNALS OF COMMON NMR SOLVENTS.....	54
TABLE 22. <sup>1</sup> H NMR CHEMICAL SHIFTS OF TRACE IMPURITIES OF COMMON SOLVENTS AND REAGENTS.....	54
TABLE 23. <sup>13</sup> C NMR CHEMICAL SHIFTS OF TRACE IMPURITIES OF COMMON SOLVENTS AND REAGENTS.....	55

TABLE 24.	AQUEOUS BUFFERS .....	56
TABLE 25.	SOLUBILITY COEFFICIENTS OF AIR AT 1atm IN WATER ....	57
TABLE 26.	SOLUBILITY COEFFICIENTS OF O <sub>2</sub> AT 1atm IN WATER .....	57
TABLE 27.	BUNSEN COEFFICIENTS ( $\beta$ ) OF GASES AT 1atm IN ORGANIC SOLVENTS AT 20°C .....	57
TABLE 28.	OSTWALD COEFFICIENTS ( <i>I</i> )/L OF O <sub>2</sub> AT 1atm IN AQUEOUS SOLUTIONS AT 25°C .....	58
TABLE 29.	SOLUBILITIES OF HCl AND NH <sub>3</sub> AT 760mm (g/100g OF SOLUTION) .....	59
TABLE 30.	BOILING POINTS OF SOME USEFUL GASES AT 760 mm .....	59
TABLE 31.	PREFIXES FOR QUANTITIES .....	59
BIBLIOGRAPHY .....		60

## CHAPTER 2

### CHEMICAL METHODS USED IN PURIFICATION

GENERAL REMARKS.....	71
REMOVAL OF TRACES OF METALS FROM REAGENTS .....	71
METAL IMPURITIES .....	71
DISTILLATION .....	71
SCAVENGER RESINS AND OTHER SUPPORTS .....	72
USE OF ION-EXCHANGE RESINS.....	72
PRECIPITATION .....	72
Removal of lead contaminants .....	72
Removal of iron contaminants.....	72
Removal of other metal contaminants.....	73
EXTRACTION .....	73
COMPLEXATION .....	73
USE OF METAL HYDRIDES .....	73
LITHIUM ALUMINIUM HYDRIDE .....	73
CALCIUM HYDRIDE .....	73
SODIUM BOROHYDRIDE.....	74
POTASSIUM BOROHYDRIDE.....	74
PURIFICATION <i>via</i> DERIVATIVES.....	74
ALCOHOLS .....	74
ALDEHYDES.....	75
AMINES .....	75

Picrates.....	75
Salts.....	76
Double salts.....	76
<i>N</i> -Acetyl derivatives.....	76
<i>N</i> -Tosyl derivatives.....	76
AROMATIC HYDROCARBONS.....	76
Adducts.....	76
Sulfonation.....	76
CARBOXYLIC ACIDS.....	77
4-Bromophenacyl esters.....	77
Alkyl esters.....	77
Salts.....	77
HYDROPEROXIDES.....	77
KETONES.....	77
Bisulfite adduct.....	77
Semicarbazones.....	77
PHENOLS.....	78
Benzoates.....	78
Acetates.....	78
PHOSPHATE AND PHOSPHONATE ESTERS.....	78
MISCELLANEOUS.....	78
GENERAL METHODS FOR THE PURIFICATION OF CLASSES OF COMPOUNDS..	78
PROCEDURES.....	78
CRITERIA OF PURITY.....	79
GENERAL PROCEDURES FOR THE PURIFICATION OF SOME CLASSES OF	
ORGANIC COMPOUNDS.....	80
ACETALS.....	80
ACIDS.....	80
Carboxylic acids.....	80
Sulfonic acids.....	81
Sulfinic acids.....	81
ACID CHLORIDES.....	81
ALCOHOLS.....	81
Monohydric alcohols.....	81
Polyhydric alcohols.....	82
ALDEHYDES.....	82
AMIDES.....	82
AMINES.....	82
AMINO ACIDS.....	83
ANHYDRIDES.....	83
CAROTENOIDS.....	83
ESTERS.....	83

ETHERS.....	84
HALIDES.....	84
HYDROCARBONS.....	85
IMIDES.....	86
IMINO COMPOUNDS.....	86
KETONES.....	86
MACROMOLECULES.....	86
NITRILES.....	86
NITRO COMPOUNDS.....	86
NUCLEIC ACIDS.....	87
PHENOLS.....	87
POLYPEPTIDES AND PROTEINS.....	87
QUINONES.....	87
SALTS.....	87
With metal ions.....	87
With organic cations.....	87
With sodium alkane sulfonates.....	88
SULFUR COMPOUNDS.....	88
Disulfides.....	88
Sulfones.....	88
Sulfoxides.....	88
Thioethers (sulfides).....	88
Thiols.....	88
Thiolsulfonates (disulfoxides).....	88
PURIFICATION <i>via</i> FLUOROCHROMES.....	88
 BIBLIOGRAPHY.....	 89

## CHAPTER 3

### THE FUTURE OF PURIFICATION

INTRODUCTION.....	91
ATOM ECONOMY.....	92
ORGANOCATALYSIS.....	92
MICROWAVE TECHNOLOGIES.....	93
FLOW CHEMISTRY.....	93
SOLID PHASE SYNTHESIS.....	94

SOLID PHASE PEPTIDE SYNTHESIS (SPPS).....	94
SOLID PHASE DEOXYRIBONUCLEOTIDE SYNTHESIS.....	95
POLYMER SUPPORTED REACTANTS.....	95
SCAVENGER RESINS.....	95
COMBINATORIAL CHEMISTRY.....	95
ALTERNATIVE SOLVENTS.....	96
SUPERCritical FLUIDS.....	96
IONIC LIQUIDS.....	96
FLUOROUS CHEMISTRY.....	97
BIOMASS DERIVED ORGANIC SOLVENTS.....	97
BIBLIOGRAPHY.....	98

## CHAPTER 4

### PURIFICATION OF ORGANIC CHEMICALS

INTRODUCTION.....	103
ALIPHATIC COMPOUNDS.....	104
ALICYCLIC COMPOUNDS.....	219
AROMATIC COMPOUNDS.....	258
HETEROCYCLIC COMPOUNDS.....	397
MISCELLANEOUS As, B, P, Si, S, Se and Te COMPOUNDS.....	499

## CHAPTER 5

### PURIFICATION OF INORGANIC AND METAL-ORGANIC CHEMICALS

INTRODUCTION.....	555
INORGANIC COMPOUNDS.....	556
METAL-ORGANIC COMPOUNDS.....	626

## CHAPTER 6

### CATALYSTS

INTRODUCTION.....	662
CATALYSTS—PART 1.....	663

HETEROGENEOUS METAL CATALYSTS.....	663
HOMOGENEOUS METAL CATALYSTS .....	671
ORGANOCATALYSTS.....	712
CATALYSTS—PART 2.....	721
LIGANDS & REAGENTS USED FOR MAKING LIGANDS THAT ASSIST CATALYSIS.....	721
PHASE TRANSFER CATALYSTS.....	751
IMIDAZOLIUM IONIC LIQUID CRYSTAL CATALYSTS.....	755
CHIRAL AUXILIARIES.....	757
LEWIS AND BRØNSTED/LOWRY ACIDS AND BASES .....	760
BIBLIOGRAPHY .....	761

## CHAPTER 7

### PURIFICATION OF BIOCHEMICALS AND RELATED PRODUCTS

INTRODUCTION.....	765
AMINO ACIDS and PEPTIDES .....	772
PROTEINS, ENZYMES, DNA and RNA .....	798
CARBOHYDRATES.....	816
CAROTENOIDS.....	835
STEROIDS.....	842
MISCELLANEOUS COMPOUNDS (including biologically useful reagents, Low-molecular-weight bioactive substances, antibiotics, coenzymes, vitamins, lipids, phospholipids, nucleosides, nucleotides and polynucleotides) .....	857

## CHAPTER 8

### NANOMATERIALS AND NANOTECHNOLOGY

INTRODUCTION.....	908
IDENTIFICATION AND MEASUREMENT OF NANOMATERIALS .....	908
FULLERENES AND RELATED SUBSTANCES.....	909
FULLERENES .....	909
FUNCTIONALISED FULLERENES.....	913
CARBON NANOTUBES (CNTs) AND RELATED MATERIALS .....	919
PREPARATION OF CNTs (NTs).....	920
PURIFICATION OF CNTs (NTs) .....	921
PREPARATION AND PURIFICATION OF FULLERENE PIPES (TUBES).....	922
SOLUBILISATION OF CNTs.....	923
CARBON NANOTUBES (CNTs) .....	924
FUNCTIONALISED CNTs .....	925

GRAPHENE MATERIALS.....	929
SELF ASSEMBLED MONOLAYERS (SAMs).....	931
DIAMOND NANOMATERIALS.....	932
SOLVENT RESISTANT NANO FILTERS (SRNF).....	933
NANO METALS AND METAL DERIVATIVES.....	933
SOME ORGANIC AND METAL-ORGANIC COMPOUNDS USED IN NANOTECHNOLOGY.....	938
SAFETY ISSUES.....	939
BIBLIOGRAPHY.....	941
<b>GENERAL INDEX.....</b>	<b>943</b>
<b>CAS REGISTRY NUMBERS INDEX.....</b>	<b>959</b>
HOW TO VALIDATE CHEMICAL ABSTRACTS SERVICE REGISTRY NUMBERS	959

## HOW THIS BOOK SHOULD BE USED

Substances have been entered under their respective chapters, sections and subsections. In these sections compounds have been entered in alphabetical order according to the more commonly used name. However, because compounds can be named differently (some of these other names have been included in brackets after their entries) and may be difficult to find, it is advisable to obtain the page number of the entry from the Chemical Abstracts Registry Numbers (CASRN) Index at the end of the book. CASRNs of substances are readily obtained from "SciFinder", or better, from any commercial catalogue that sells these compounds, as almost all of these have CASRNs inserted after the names of their products. Also, in this book we could not insert some substances that can be formally included in more than one section without repeating the entry. In such cases the compounds are entered in a preferred section and are cross-referenced by inserting "see [CASRN]" or just "[CASRN]". Thus the cross-reference is its CASRN. The CASRNs Index provides the page numbers in bold/italic type. If the CASRN of a desired substance is not in the CASRN Index, or its full or abbreviated name is not in the General Index, then it will not be present in this book.

CASRNs are unique for each chemical substance and are such that they are internally consistent. They are set up according to a specific formula. Refer to the first page of the CASRNs Index to calculate the formula in order to check whether the number is a valid number or not.

W.L.F.A. & C.L.L.C.

## ABOUT THE AUTHORS

### **W.L.F. Armarego**

Division of Molecular Bioscience, The John Curtin School of Medical Research, Australian National University, Canberra

Wilfred L. F. Armarego graduated BSc (Hons) in 1953 and PhD from the University of London in 1956 and came to Australia in that year. After two years at the Central Research Laboratories (ICIANZ) in Melbourne, where he worked on plant growth substances, and one year on potentially carcinogenic polycyclic aromatic hydrocarbons at the University of Melbourne as Senior Demonstrator in Organic Chemistry, he joined the Department of Medical Chemistry as a Research Fellow in 1960. He became a Fellow in 1963 and was awarded a DSc degree (London) in 1968. He was promoted to Senior Fellow in 1967 and began research work on the biochemistry and molecular biology of pteridine-requiring enzymes related to the inherited metabolic disease phenylketonuria and its variants. He was head of the Protein Biochemistry Group and Pteridine Biochemistry Laboratory and is now a visiting fellow at the John Curtin School of Medical Research, and member of the editorial boards of 'Medicinal Research Reviews' and 'Pteridines' journals. He is a Chartered Chemist, a Fellow of the Royal Society of Chemistry and a Fellow of the Royal Australian Chemical Institute.

### **Christina Chai**

Department of Pharmacy, National University of Singapore and Institute of Chemical and Engineering Sciences, Agency for Science Technology and Research, Singapore.

Christina Li Lin Chai obtained her BSc (Hons) from the University of Canterbury, Christchurch, New Zealand and her PhD in synthetic organic chemistry from the Research School of Chemistry, Australian National University, Canberra, Australia. Upon completion of her PhD, she was awarded a Samuel-Glasstone Research Fellowship tenable at the Dyson Perrins Laboratory, University of Oxford, UK. Her first faculty position was at the Department of Chemistry, Victoria University of Wellington in 1991-1993, followed by a second faculty position at the Australian National University (1994-2004) where she was an Associate Professor. She moved to the Institute of Chemical and Engineering Sciences, Agency for Science Technology and Research (A\*STAR), Singapore as a Principal Scientist and Programme Manager in Dec 2004 and held various co-appointments as Director of Graduate Affairs at the Science and Engineering Council (SERC), A\*STAR as well as an honorary adjunct appointment at Nanyang Technological University, Singapore. Since August 2011, Dr Chai holds a joint appointment as Associate Professor in the Department of Pharmacy, National University of Singapore and as Principal Scientist in the Institute of Chemical and Engineering Sciences.

# CHAPTER 1

## COMMON PHYSICAL TECHNIQUES

### USED IN PURIFICATION

#### INTRODUCTION

Purity is a matter of degree. Other than contaminants such as dust, paper fibres, wax, cork, etc., that may have been inadvertently introduced into the sample during manufacture, all commercially available chemical substances are in some measure impure. Any amounts of unreacted starting material, intermediates, by-products, isomers and related compounds may be present depending on the synthetic or isolation procedures used for preparing the substances. Inorganic reagents may deteriorate because of defective packaging (glued liners affected by sulfuric acid, zinc extracted from white rubber stoppers by ammonia), corrosion or prolonged storage. Organic molecules may undergo changes on storage. In extreme cases the container may be incorrectly labelled or, where compositions are given, they may be misleading or inaccurate for the proposed use. Where any doubt exists, it is usual to check for impurities by appropriate spot tests, or by recourse to tables of physical or spectral properties such as the extensive infrared and NMR libraries published by the Sigma-Aldrich Chemical Co.

The important question, then, is not whether a substance is pure but whether a given sample is sufficiently pure for some intended purpose. That is, are the contaminants likely to interfere in the process or measurement that is to be studied? By suitable manipulation it is often possible to reduce levels of impurities to acceptable limits, but absolute purity is an ideal which, no matter how closely approached, can never be attained. A *negative* physical or chemical test indicates only that the amount of an impurity in a substance lies below a certain sensitivity level; no test can demonstrate that a likely impurity is entirely absent.

When setting out to purify a laboratory chemical, it is desirable that the starting material is of the best grade commercially available. Particularly among organic solvents there is a range of qualities varying from *laboratory chemical* to *spectroscopic* and *chromatographic* grades. Many of these are suitable for use as received. With the more common reagents it is usually possible to obtain from the current literature some indications of likely impurities, their probable concentrations and methods for detecting them. However, in many cases complete analyses are not given so that significant concentrations of unspecified impurities may be present.

#### THE QUESTION OF PURITY

Solvents and substances that are specified as *pure* for a particular purpose may, in fact, be quite impure for other uses. Absolute ethanol may contain traces of benzene, which makes it unsuitable for ultraviolet spectroscopy, or plasticizers which make it unsuitable for use in solvent extraction. See also the section on "Criteria of Purity" in Chapter 2.

Irrespective of the grade of material to be purified, it is essential that some criteria exist for assessing the degree of purity of the final product. The more common of these include:

1. Examination of physical properties such as:
  - (a) Melting point, freezing point, boiling point, and the freezing curve (i.e. the variation, with time, in the freezing point of a substance that is being slowly and continuously frozen).
  - (b) Density.

- (c) Refractive index at a specified temperature and wavelength. The sodium D line at 589.26 nm (weighted mean of the D<sub>1</sub> and D<sub>2</sub> lines) is the usual wavelength used but the refractive index values at other wavelengths can often be interpolated from a plot of refractive index versus 1/(wavelength)<sup>2</sup>.
- (d) Specific conductivity (This can be used to detect, for example, water, salts, inorganic and organic acids and bases, in non-electrolytes).
- (e) Optical rotation, optical rotatory dispersion and circular dichroism.

2. Empirical analysis, for C, H, N, ash, etc.

3. Chemical tests for particular types of impurities, e.g. for peroxides in aliphatic ethers (with acidified KI), or for water in solvents (quantitatively by the Karl Fischer method, see Fieser and Fieser, *Reagents for Organic Synthesis*, J. Wiley & Sons, NY, Vol 1 pp. 353, 528 1967, Library of Congress Catalog Card No 66-27894, also see Karl Fischer titrant or Hydranal<sup>®</sup> –Titrant type 5E [64-17-5] and other types in Fluka and Sigma-Aldrich Catalogue.

4. Physical tests for particular types of impurities:

Emission and atomic absorption spectroscopy for detecting organic impurities and determining metal ions.

Chromatography, including paper, thin layer, liquid (high, medium and normal pressure), flash and vapour phase.

Electron spin resonance for detecting free radicals.

Other spectroscopic methods (see 5 below).

5. Examination of spectroscopic properties

Nuclear Magnetic Resonance (<sup>1</sup>H, <sup>13</sup>C, <sup>31</sup>P, <sup>19</sup>F, <sup>11</sup>B NMR, etc.)

Infrared spectroscopy (IR and Fourier Transform IR)

Ultraviolet (UV), visible and fluorescence spectroscopy

X-ray photoelectron spectroscopy (XPS)

Atomic absorption spectroscopy (AAA)

Mass spectroscopy [electron ionisation (EI), chemical ionisation (CI), electrospray ionisation (ESI), fast atom bombardment (FAB), matrix-associated laser desorption ionisation (MALDI), inductively coupled plasma-mass spectrometry (ICP-MS)]

6. Electrochemical methods (see Chapter 6 for macromolecules).

7. Nuclear methods which include a variety of radioactive elements as in organic reagents, complexes or salts.

A substance is usually taken to be of an acceptable purity when the measured property is unchanged by further treatment (especially if it agrees with a recorded value). In general, at least two different methods, such as recrystallisation and distillation, should be used in order to ensure maximum purity. Crystallisation may be repeated (from the same solvent or better from different solvents) until the substance has a constant melting point, and until it distils repeatedly within a narrow specified temperature range. The purified product should have spectroscopic properties which indicate that the traces of impurities left in the sample are of acceptable levels for the intended purpose.

With liquids, the refractive index at a specified temperature and wavelength is a sensitive test of purity. Note however that this is sensitive to dissolved gases such as O<sub>2</sub>, N<sub>2</sub> or CO<sub>2</sub>. Under favourable conditions, freezing curve studies are sensitive to impurity levels of as little as 0.001 moles percent. Analogous fusion curves or heat capacity measurements can be up to ten times as sensitive as this. With these exceptions, most of the above methods are rather insensitive, especially if the impurities and the substances in which they occur are chemically similar. In some cases, even an impurity comprising many parts per million of a sample may escape detection.

The common methods of purification, discussed below, comprise distillation (including fractional distillation, distillation under reduced pressure, sublimation and steam distillation), crystallisation, extraction, chromatographic, electrophoresis and other methods. In some cases, volatile and other impurities can be removed simply by heating. Impurities can also sometimes be eliminated by the formation of derivatives from which the purified material is regenerated (see Chapter 2).

Common techniques used for determining the purity of a specimen and for following the purification procedure are described below (p. 33).

### SOURCES OF IMPURITIES

Some of the more obvious sources of contamination of solvents arise from storage in metal drums and plastic containers, and from contact with grease and screw caps. Many solvents contain water. Others have traces of acidic materials such as hydrochloric acid in chloroform. In both cases this leads to corrosion of the drum and contamination of the solvent by traces of metal ions, especially  $\text{Fe}^{3+}$ . Grease, for example on stopcocks of separating funnels and other apparatus, e.g. greased ground joints, is also likely to contaminate solvents during extractions and chemical manipulation. Oxygen from the air is also a source of contamination by virtue of its ability to produce small or large amounts of oxidation products (see section on the *Solubilities of gases in liquids* below).

A much more general source of contamination that has not received the consideration it merits comes from the use of plastics for tubing and containers. Plasticisers can readily be extracted by organic solvents from PVC and other plastics, so that most solvents, irrespective of their grade (including spectrograde and ultrapure), have been reported to contain 0.1 to 5ppm of plasticiser [de Zeeuw, Jonkman and van Mansvelt *Anal Biochem* **67** 339 1975]. Where large quantities of solvent are used for extraction followed by evaporation, this can introduce significant amounts of impurity, even exceeding the weight of the genuine extract and giving rise to spurious peaks in gas chromatography, for example of fatty acid methyl esters [Pascaud, *Anal Biochem* **18** 570 1967]. Likely contaminants are di(2-ethylhexyl)phthalate and dibutyl phthalate, but upwards of 20 different phthalate esters are listed as plasticisers as well as adipates, azelates, phosphates, epoxides, polyesters and various heterocyclic compounds. These plasticisers would enter the solvent during passage through plastic tubing or from storage in containers or from plastic coatings used in cap liners for bottles. Such contamination could arise at any point in the manufacture or distribution of a solvent. The problem with cap liners is avoidable by using corks wrapped in aluminium foil, although even in this case care should be taken because aluminium foil can dissolve in some liquids, e.g. benzylamine and propionic acid. Polycarbonate containers invariably leach out the 'estrogenic chemical' **Bisphenol A** (see Chapter 4, Aromatic Compounds) into the liquid in the container [Fiona Case *Chemistry World* **5** (No. 4) 12 2008, Rebecca Trager *Chemistry World* **5** (No. 5) 8 2008].

Solutions in contact with polyvinyl chloride can become contaminated with trace amounts of lead, titanium, tin, zinc, iron, magnesium or cadmium from additives used in the manufacture and moulding of PVC.

*N*-Phenyl-2-naphthylamine is a contaminant of solvents and biological materials that have been in contact with black rubber or neoprene (in which it is used as an antioxidant). Although this naphthylamine was only an artifact of the isolation procedures, it was at first thought to be a genuine component of vitamin K preparations, extracts of plant lipids, algae, butter, animal livers, eye tissue and kidney tissue [Brown *Chem Br* **3** 524 1967].

Most of the above impurities can be removed by prior distillation of the solvent, and care should be taken to avoid further contact with plastic or black rubber materials.

### PRACTICES TO AVOID IMPURITIES

#### Cleaning practices

Laboratory glassware and Teflon equipment can be cleaned satisfactorily for most purposes by careful immersion into a solution of sodium dichromate in concentrated sulfuric acid, followed by draining, and rinsing copiously with distilled water. This is an exothermic reaction and should be carried out **very** cautiously in an efficient fume cupboard. [To prepare the chromic acid bath, dissolve 5 g of sodium dichromate (CARE: cancer suspect agent) in 5 ml of water. The dichromate solution is then cooled and stirred while 100 ml of concentrated sulfuric acid is added slowly. Store it in a glass bottle.] Where traces of chromium (adsorbed on the glass) must be avoided, a 1:1 mixture of concentrated sulfuric and nitric acid is a useful alternative. (*Use in a fumehood to remove vapour and with adequate face protection.*) Acid washing is also suitable for polyethylene ware, but prolonged contact (some weeks) leads to severe deterioration of the plastic. Alternatively an alcoholic solution of sodium hydroxide (alkaline base bath) can be used. This strongly corrosive solution (CAUTION: alkali causes serious burns) can be made by dissolving 120 g of NaOH in 120 ml of water, followed by dilution to 1 L with 95% ethanol.

This solution is conveniently stored in suitable alkali-resistant containers (e.g. Nalgene heavy duty rectangular tanks) with lids. Glassware can be soaked overnight in the base bath and rinsed thoroughly after soaking. For much glassware, washing with hot detergent solution, using tap water, followed by rinsing well with distilled water and acetone, and heating to 200–300° overnight, is adequate. (Volumetric apparatus should not be heated: after washing it is rinsed with acetone, then pure diethyl ether, and air-dried. Prior to use, equipment can be rinsed with acetone, then with petroleum ether or pure diethyl ether, to remove the last traces of contaminants.) Teflon equipment should be soaked, first in acetone, then in petroleum ether or pure diethyl ether for ten minutes, then dried in a vacuum or flushed with dry nitrogen prior to use.

For trace metal analyses, prolonged soaking of equipment in 1M nitric acid may be needed to remove adsorbed metal ions.

Soxhlet thimbles and filter papers may contain traces of lipid-like materials. For manipulations with highly pure materials, as in trace-pesticide analysis, thimbles and filter papers should be thoroughly extracted with pure diethyl ether before use.

Trace impurities in silica gel for TLC can be removed by heating at 300° for 16 hours or by Soxhlet extraction for 3 hours with distilled chloroform, followed by 4 hours extraction with distilled pure diethyl ether and drying in a vacuum.

### Silylation of glassware and plasticware

Silylation of apparatus makes it repellant to water and hydrophilic materials. It minimises loss of solute by adsorption onto the walls of the container. The glassware is placed in a desiccator containing dichloromethyl silane (1ml) in a small beaker and evacuated for 5 minutes. The vacuum is turned off and air is introduced into the desiccator, which allows the silylating agent to coat the glassware uniformly. The desiccator is then evacuated, closed and set aside for 2 hours. The glassware is removed from the desiccator and baked at 180° for 2 hours before use.

Fluka supplies a variety of silylating mixtures including (a) Silanization solution I (Fluka gel repel I: ~5% dimethyldichlorosilane in hexane, (b) Silanization solution II (Fluka gel repel II: ~2% dimethyldichlorosilane in 1,1,1-trichloroethane) for silanizing micro electrode, (c) Silanization solution III (Selectophore: 10% hexamethyldisilazane and 6% trimethylchlorosilane in 1-chloronaphthalene), (d) Silanization solution IV (Selectophore: ~4% trimethylchlorosilane in *o*-xylene, (e) Silanization solution V (Selectophore: ~5% dimethyldichlorosilane in *o*-xylene, and (f) Silanization solution VI (Selectophore: ~3% tributylchlorosilane in 1-chloronaphthalene for silanizing micropipette electrodes by the dip-and-bake method. Most powerful general silylating mixtures among many others are also available from Fluka (see catalogue).

Plasticware is treated similarly except that it is rinsed well with water before use instead of baking. Note that dichloromethyl silane is highly **TOXIC** and **VOLATILE**, and the whole operation should be carried out in an efficient fume cupboard.

An alternative procedure used for large apparatus is to rinse the apparatus with a 5% solution of dichloromethyl silane in chloroform, followed by several rinses with water before baking the apparatus at 180°/2 hours (for glass) or drying in air (for plasticware).

A solution of 2% w/v of dichloromethyl silane in octamethyl cyclooctasilane or octmethylcyclotetrasiloxane is used to inhibit the sticking of polyacrylamide gels, agarose gels and nucleic acids to glass surfaces and these chemicals are available commercially [from Fluka (Riedel-deHaën)].

## SAFETY PRECAUTIONS ASSOCIATED WITH THE PURIFICATION OF LABORATORY CHEMICALS

Although most of the manipulations involved in purifying laboratory chemicals are inherently safe, care is necessary if hazards are to be avoided in the chemical laboratory. In particular there are dangers inherent in the inhalation of vapours and absorption of liquids and low melting solids through the skin. In addition to the toxicity of solvents there is also the risk of their flammability and the possibility of eye damage. Chemicals, particularly in admixture, may be explosive. Compounds may be carcinogenic or otherwise deleterious to health. Present-day chemical catalogues specifically indicate the particular dangerous properties of the individual chemicals they list, and these should be consulted whenever the use of commercially available chemicals is contemplated. Radioisotopic labeled compounds pose special problems of human exposure and of disposal of laboratory waste.

Hazardous purchased chemicals are accompanied by detailed MSDS (Material Safety Data Sheets), which contain information regarding their toxicity, safety handling procedures and the necessary precautions to be taken. These should be read carefully and filed for future reference. In addition, chemical management systems such as Chem-ChemWatch, which include information on hazards, handling and storage, are commercially available. There are a number of websites which provide selected safety information: these include the Sigma-Aldrich website <[www.sigmaaldrich.com](http://www.sigmaaldrich.com)> and other chemical websites, e.g. <[www.ilpi.com/msds](http://www.ilpi.com/msds)>.

The most common hazards are:

- (1) Explosions due to the presence of peroxides formed by aerial oxidation of ethers and tetrahydrofuran, decahydronaphthalene, acrylonitrile, styrene and related compounds.
- (2) Compounds with low flash points (below room temperature). Examples are acetaldehyde, acetone, acetonitrile, benzene, carbon disulfide, cyclohexane, diethyl ether, ethyl acetate and *n*-hexane.
- (3) Contact of oxidising agents (KMnO<sub>4</sub>, HClO<sub>4</sub>, chromic acid) with organic liquids.
- (4) Toxic reactions with tissues (Me<sub>2</sub>SO<sub>4</sub>).

The laboratory should at least be well-ventilated and safety glasses should be worn, particularly during distillations and manipulations carried out under reduced pressure or elevated temperatures. With this in mind we have endeavoured to warn users of this book whenever greater than usual care is needed in handling chemicals. As a general rule, however, **all chemicals which users are unfamiliar with should be treated with extreme care and assumed to be highly flammable and toxic.** The safety of others in a laboratory should always be foremost in mind, with ample warning whenever a potentially hazardous operation is in progress. Also, unwanted solutions or solvents should never be disposed of *via* the laboratory sink. The operator should be aware of the usual means for disposal of chemicals in her/his laboratories, and she/he should remove unwanted chemicals accordingly. **Organic liquids for disposal should be temporarily stored, as is practically possible, in respective containers. Avoid placing all organic liquids in the same container particularly if they contain small amounts of reagents which could react with each other. Halogenated waste solvents should be kept separate from other organic liquids.**

**Laboratory coats, disposable aprons, caps, sleeves, dust/mist respirators and foot protection, hearing protection as well as a variety of safety glasses, goggles, face and body shields should be used when the demand arises and are available commercially (see e.g. the Sigma-Aldrich Labware catalogue).**

## SOME HAZARDS OF CHEMICAL MANIPULATION IN PURIFICATION AND RECOVERY OF RESIDUES

Performing chemical manipulations calls for some practical knowledge if danger is to be avoided. However, with care, hazards can be kept to an acceptable minimum. A good general approach is to consider every operation as potentially perilous and then to adjust one's attitude as the operation proceeds. A few of the most common dangers are set out below. For a larger coverage of the following sections, and of the literature, the bibliography at the end of this chapter should be consulted.

**Perchlorates and perchloric acid.** At 160° perchloric acid is an exceedingly strong oxidising acid and a strong dehydrating agent. Organic perchlorates, such as methyl and ethyl perchlorates, are unstable and are violently explosive compounds. A number of heavy-metal perchlorates are extremely prone to explode. The use of anhydrous magnesium perchlorate, *Anhydrone*, *Dehydrite*, as a drying agent for organic vapours is **not** recommended. Desiccators which contain this drying agent should be adequately shielded at all times and kept in a cool place, i.e. **never** on a window sill where sunlight can fall on it.

No attempt should be made to purify perchlorates, except for ammonium, alkali metal and alkaline earth salts which, in water or aqueous alcoholic solutions are insensitive to heat or shock. Note that perchlorates react relatively slowly in aqueous organic solvents, but as the water is removed there is an increased possibility of an explosion. Perchlorates, often used in non-aqueous solvents, are explosive in the presence of even small amounts of organic compounds when heated. Hence stringent care should be taken when purifying perchlorates, and direct flame and infrared lamps should be avoided. Tetra-alkylammonium perchlorates should be dried below 50° under vacuum (and protection). Only very small amounts of such materials should be prepared, and stored, at any one time.

**Peroxides.** These are formed by aerial oxidation or by autoxidation of a wide range of organic compounds, including diethyl ether, allyl ethyl ether, allyl phenyl ether, dibenzyl ether, benzyl butyl ether, *n*-butyl ether, *iso*-butyl ether, *t*-butyl ether, dioxane, tetrahydrofuran, olefins, and aromatic and saturated aliphatic hydrocarbons. They accumulate during distillation and can detonate violently on evaporation or distillation when their concentration becomes high. If peroxides are likely to be present materials should be tested for peroxides before distillation (for tests see entry under "Ethers", in Chapter 2). Also, distillation should be discontinued when at least one quarter of the residue is left in the distilling flask.

**Heavy-metal-containing explosives.** Ammoniacal silver nitrate, on storage or treatment, will eventually deposit the highly explosive silver nitride *fulminating silver*. Silver nitrate and ethanol may give silver fulminate (see Chapter 5), and in contact with azides or hydrazine and hydrazides may form silver azide. Mercury can also form such compounds. Similarly, ammonia or ammonium ions can react with gold salts to form "*fulminating gold*". Metal fulminates of cadmium, copper, mercury and thallium are powerfully explosive, and some are detonators [Luchs, *Photog Sci Eng* **10** 334 1966]. Heavy-metal-containing solutions, particularly when organic material is present, should be treated with great respect and precautions towards possible explosion should be taken.

**Strong acids.** In addition to perchloric acid (see above), extra care should be taken when using strong mineral acids. Although the effects of concentrated sulfuric acid are well known, these cannot be stressed strongly enough. Contact with tissues will leave irreparable damage. **Always dilute the concentrated acid by carefully adding the acid down the side of the flask which contains the water, and the process should be carried out under cooling. This solution is not safe to handle until the acid has been thoroughly mixed (care) with the water. Protective face, and body coverage should be used at all times.** Fuming sulfuric acid and chlorosulfonic acid are even more dangerous than concentrated sulfuric acid, and adequate precautions should be taken. Chromic acid cleaning mixture contains strong sulfuric acid and should be treated in the same way; and in addition the mixture is potentially *carcinogenic*. Concentrated and fuming nitric acids are also dangerous because of their severe deleterious effects on tissues.

**Picric acid.** This acid and related nitro compounds, e.g. styphnic acid, are explosive and should **not** be kept dry. The acid is generally stored wet by covering the crystals with water. Solutions in ethanol and benzene are used occasionally. They should be stored in the cold (to minimise evaporation), and a rubber or plastic stopper (**not a ground glass stopper**) should be used. **Note** that picric acid and picrates stain skin protein with a yellow colour that is not readily washed off. This can be avoided by wearing rubber gloves.

**Reactive halides and anhydrides.** Substances like acid chlorides, low-molecular-weight anhydrides and some inorganic halides (e.g.  $\text{PCl}_3$ ) can be **highly toxic and lachrymatory, affecting mucous membranes and lung tissues. Utmost care should be taken when working with these materials. Work should be carried out in a very efficient fume cupboard.**

**Salts and organic esters of some inorganic acids.** In addition to the dangers of perchlorate salts, other salts such as nitrates, azides, diazo salts, organic nitrates, organic azides and picrates (see above) can be hazardous, and due care should be taken when these are dried. Large quantities should never be prepared or stored for long periods.

**Solvents.** The flammability of low-boiling organic liquids cannot be emphasised strongly enough. These invariably have very low flash points and can ignite spontaneously. Special precautions against explosive flammability should be taken when recovering such liquids. Care should be taken with small volumes (*ca* 250ml) as well as large volumes (> 1L), and the location of all the fire extinguishers, and fire blankets, in the immediate vicinity of the apparatus should be checked. The fire extinguisher should be operational. The following flammable liquids (in alphabetical order) are common fire hazards in the laboratory: acetaldehyde, acetone, acrylonitrile, acetonitrile, benzene, carbon disulfide, cyclohexane, diethyl ether, ethyl acetate, hexane, low-boiling petroleum ether, tetrahydrofuran and toluene. Toluene should always be used in place of benzene wherever possible due to the potential *carcinogenic* effects of the liquid and vapour of the latter. The drying of flammable solvents with sodium or potassium metal and metal hydrides poses serious potential fire hazards, and adequate precautions should be stressed.

## SAFETY DISCLAIMER

Experimental chemistry is a very dangerous occupation, and extreme care and adequate safety precautions should be taken at all times. Although we have stated the safety measures that have to be taken under specific entries, these are by no means exhaustive and some may have been unknowingly or accidentally omitted. The experimenter without prior knowledge or experience must seek further safety advice on reagents and procedures from experts in the field before undertaking the purification of any material. *We take no responsibility whatsoever if any mishaps occur when using any of the procedures described in this book.*

## METHODS OF PURIFICATION OF REAGENTS AND SOLVENTS

Many methods exist for the purification of reagents and solvents. A number of these methods are routinely used in synthetic as well as analytical chemistry and biochemistry. These techniques, outlined below, will be discussed in greater detail in the respective sections in this chapter. It is important to note that more than one method of purification may need to be implemented in order to obtain compounds of highest purity.

Common methods of purification are:

- Solvent Extraction and Distribution
- Distillation
- Recrystallisation
- Sublimation
- Electrophoresis
- Chromatography

For substances contaminated with water or solvents, drying with appropriate absorbents and desiccants may be sufficient.

## SOLVENT EXTRACTION AND DISTRIBUTION

Extraction of a substance from suspension or solution into another solvent can sometimes be used as a purification process. Thus, organic substances can often be separated from inorganic impurities by shaking an aqueous solution or suspension with suitable immiscible solvents such as benzene, carbon tetrachloride, chloroform, diethyl ether, diisopropyl ether or petroleum ether. After several such extractions, the combined organic phase is dried and the solvent is evaporated. Grease from the glass taps of conventional separating funnels is invariably soluble in the solvents used. Contamination with grease can be very troublesome particularly when the amounts of material to be extracted are very small. Instead, the glass taps should be lubricated with the extraction solvent; or better, the taps of the extraction funnels should be made of the more expensive material *Teflon*. Immiscible solvents suitable for extractions are given in Table 1. Addition of electrolytes (such as ammonium sulfate, calcium chloride or sodium chloride) to the aqueous phase helps to ensure that the organic layer separates cleanly and also decreases the extent of extraction into the latter. Emulsions can also be broken up by filtration (with suction) through Celite, or by adding a little diethyl ether, octyl alcohol or some other paraffinic alcohol. The main factor in selecting a suitable immiscible solvent is to find one in which the material to be extracted is readily soluble, whereas the substance from which it is being extracted is not. The same considerations apply irrespective of whether it is the substance being purified, or one of its contaminants, that is taken into the new phase. (The second of these processes is described as washing.)

Common examples of washing with aqueous solutions include the following:

- Removal of acids from water-immiscible solvents by washing with aqueous alkali, sodium carbonate or sodium bicarbonate.
- Removal of phenols from similar solutions by washing with aqueous alkali.
- Removal of organic bases by washing with dilute hydrochloric or sulfuric acids.
- Removal of unsaturated hydrocarbons, of alcohols and of ethers from saturated hydrocarbons or alkyl halides by washing with cold concentrated sulfuric acid.

This process can also be applied to purification of the substance if it is an acid, a phenol or a base, by extracting into the appropriate aqueous solution to form the salt which, after washing with pure solvent, is again converted to the free species and re-extracted. Paraffin hydrocarbons can be purified by extracting them with phenol (in which aromatic hydrocarbons are highly soluble) prior to fractional distillation.

For extraction of solid materials with a solvent, a *Soxhlet* extractor is commonly used. This technique is applied, for example, in the alcohol extraction of dyes to free them from insoluble contaminants such as sodium chloride or sodium sulfate.

Acids, bases and amphoteric substances can be purified by taking advantage of their ionisation constants (see below).

The recovery of some fifty more commonly used solvents from water, other solvents, residues etc. have been discussed, together with information on their behaviour before and after use, by I.M. Smallwood in the *Solvent Recovery Handbook*, Blackwood Science Publ Ltd, 2001, ISBN 9780632056477.

## DISTILLATION

One of the most widely applicable and most commonly used methods of purification of liquids or low melting solids (especially of organic chemicals) is fractional distillation at atmospheric, or some lower, pressure. Almost without exception, this method can be assumed to be suitable for all organic liquids and most of the low-melting organic solids. For this reason it has been possible, e.g. in Chapter 4, to omit many procedures for purification of organic chemicals when only a simple fractional distillation is involved—the suitability of such a procedure is implied from the boiling point.

The boiling point of a liquid varies with the 'atmospheric' pressure to which it is exposed. A liquid boils when its vapour pressure is the same as the external pressure on its surface, its normal boiling point being the temperature at which its vapour pressure is equal to that of a standard atmosphere (760 mmHg). Lowering the external pressure lowers the boiling point. For most substances, boiling point and vapour pressure are related by an equation of the form,

$$\log p = A + B/(t + 273),$$

where  $p$  is the pressure in mmHg,  $t$  is in °C, and  $A$  and  $B$  are constants. Hence, if the boiling points at two different pressures are known, the boiling point at another pressure can be calculated from a simple plot of  $\log p$  versus  $1/(t + 273)$ . For organic molecules that are not strongly associated, this equation can be written in the form,

$$\log p = 8.586 - 5.703 (T + 273)/(t + 273)$$

where  $T$  is the boiling point in °C at 760 mmHg. Tables 2A and 2B give computed boiling points over a range of pressures. Some examples illustrate its application. Ethyl acetoacetate, **b** 180° (with decomposition) at 760 mmHg has a predicted **b** of 79° at 16 mm; the experimental value is 78°. Similarly 2,4-diaminotoluene, **b** 292° at 760mm, has a predicted **b** of 147° at 8mm; the experimental value is 148–150°. For self-associated molecules the predicted **b** are lower than the experimental values. Thus, glycerol, **b** 290° at 760mm, has a predicted **b** of 146° at 8mm: the experimental value is 182°.

Similarly an estimate of the boiling points of liquids at reduced pressure can be obtained using a nomogram (see Fig. 1).

For pressures near 760mm, the change in boiling point is given approximately by

$$\delta t = a(760 - p)(t + 273)$$

where  $a = 0.00012$  for most substances, but  $a = 0.00010$  for water, alcohols, carboxylic acids and other associated liquids, and  $a = 0.00014$  for very low-boiling substances such as nitrogen or ammonia [Crafts *Chem Ber* 20 709 1887]. When all the impurities are non-volatile, simple distillation is adequate purification. The observed boiling point remains almost constant and approximately equal to that of the pure material. Usually, however, some of the impurities are appreciably volatile, so that the boiling point progressively rises during the distillation because of the progressive enrichment of the higher-boiling components in the distillation flask. In such cases, separation is effected by fractional distillation using an efficient column. [For further reading see section on "Variation of Boiling Points with Pressure" on pp. 15–26 in *CRC—Handbook of Chemistry and Physics 88<sup>th</sup> Edition*, David R Lide (Editor-in-Chief) CRC Press, Boca Raton, Florida 2007–2008, ISBN 0849304881.]

### Techniques

The distillation apparatus consists basically of a distillation flask, usually fitted with a vertical fractionating column (which may be empty, or packed with suitable materials such as glass helices or stainless-steel wool) to which is attached a condenser leading to a receiving flask. The bulb of a thermometer projects into the vapour phase just below the region where the condenser joins the column. The distilling flask is heated so that its contents are steadily vaporised by boiling. The vapour passes up into the column where, initially, it condenses and runs back into the flask. The resulting heat transfer gradually warms the column so that there is a progressive movement of the vapour phase-liquid boundary up the column, with increasing enrichment of the more volatile component. Because of this fractionation, the vapour finally passing into the condenser (where it condenses and flows into the receiver) is commonly that of the lowest-boiling components in the system. The conditions apply until all of the low-boiling material has been distilled, whereupon distillation ceases until the column temperature is high enough to permit the next component to distil. This usually results in a temporary fall in the temperature indicated by the thermometer.

### Distillation of liquid mixtures

The principles involved in fractional distillation of liquid mixtures are complex but can be seen by considering a system which approximately obeys *Raoult's law*. (This law states that the vapour pressure of a solution at any given temperature is the sum of the vapour pressures of each component multiplied by its mole fraction in the solution.) If two substances, A and B, having vapour pressures of 600 mmHg and 360 mmHg, respectively, were mixed in a molar ratio of 2:1 (i.e. 0.666:0.333 mole ratio), the mixture would have (ideally) a vapour pressure of 520 mmHg (i.e.  $600 \times 0.666 + 360 \times 0.333$ , or  $399.6 + 119.88$  mmHg) and the vapour phase would contain 77% ( $399.6 \times 100/520$ ) of A and 23% ( $119.88 \times 100/520$ ) of B. If this phase was now condensed, the new liquid phase would, therefore, be richer in the volatile component A. Similarly, the vapour in equilibrium with this phase is still further enriched in A. Each such liquid-vapour equilibrium constitutes a "theoretical plate". The efficiency of a fractionating column is commonly expressed as the number of such plates to which it corresponds in operation. Alternatively, this information may be given in the form of the height equivalent to a theoretical plate, or HETP. The number of theoretical plates and equilibria between liquids and vapours are affected by the factors listed to achieve maximum separation by fractional distillation in the section below on techniques.

In most cases, systems deviate to a greater or lesser extent from Raoult's law, and vapour pressures may be greater or less than the values calculated. In extreme cases (e.g. azeotropes), vapour pressure-composition curves pass through maxima or minima, so that attempts at fractional distillation lead finally to the separation of a constant-boiling (azeotropic) mixture and one (but not both) of the pure species if either of the latter is present in excess.

*Elevation of the boiling point by dissolved solids.* Organic substances dissolved in organic solvents cause a rise in boiling point which is proportional to the concentration of the substance, and the extent of rise in temperature is characteristic of the solvent. The following equation applies for dilute solutions and non-associating substances:

$$\frac{M \Delta t}{c} = K$$

where M is the molecular weight of the solute,  $\Delta t$  is the elevation of boiling point in °C,  $c$  is the concentration of solute in grams for 1000 gm of solvent, and K is the *Ebullioscopic Constant* (molecular elevation of the boiling point) for the solvent. K is a fixed property (constant) for the particular solvent. This has been very useful for the determination of the molecular weights of organic substances in solution.

The efficiency of a distillation apparatus used for purification of liquids depends on the difference in boiling points of the pure material and its impurities. For example, if two components of an ideal mixture have vapour pressures in the ratio 2:1, it would be necessary to have a still with an efficiency of at least seven plates (giving an enrichment of  $2^7 = 128$ ) if the concentration of the higher-boiling component in the distillate was to be reduced to less than 1% of its initial value. For a vapour pressure ratio of 5:1, three plates would achieve as much separation.

In a fractional distillation, it is usual to reject the initial and final fractions, which are likely to be richer in the lower-boiling and higher-boiling impurities respectively. The centre fraction can be further purified by repeated fractional distillation.

To achieve maximum separation by fractional distillation:

1. The column must be flooded initially to wet the packing. For this reason it is customary to operate a still at reflux for some time before beginning the distillation.

2. The reflux ratio should be high (i.e. the ratio of drops of liquid which return to the distilling flask and the drops which distil over), so that the distillation proceeds slowly and with minimum disturbance of the equilibria in the column.
3. The hold-up of the column should not exceed one-tenth of the volume of any one component to be separated.
4. Heat loss from the column should be prevented but if the column is heated to offset this, its temperature must not exceed that of the distillate in the column.
5. Heat input to the still-pot should remain constant.
6. For distillation under reduced pressure there must be careful control of the pressure to avoid flooding or cessation of reflux.

### Types of distillation

**The distilling flask.** To minimise superheating of the liquid (due to the absence of minute air bubbles or other suitable nuclei for forming bubbles of vapour), and to prevent bumping, one or more of the following precautions should be taken:

(a) The flask is heated uniformly over a large part of its surface, either by using an electrical heating mantle or, by partial immersion in a bath above the boiling point of the liquid to be distilled (Table 3).

(b) Before heating begins, small pieces of unglazed fireclay or porcelain (porous pot, boiling chips), pumice, diatomaceous earth, or platinum wire are added to the flask. These act as sources of air bubbles.

(c) The flask may contain glass siphons or boiling tubes. The former are inverted J-shaped tubes, the end of the shorter arm being just above the surface of the liquid. The latter comprise long capillary tubes sealed above the lower end.

(d) A steady slow stream of inert gas (e.g. N<sub>2</sub>, Ar or He) is passed through the liquid.

(e) The liquid in the flask is stirred mechanically. This is especially necessary when suspended insoluble material is present.

For simple distillations a Claisen flask is often used. This flask is, essentially, a round-bottomed flask to the neck of which is joined another neck carrying a side arm. This second neck is sometimes extended so as to form a Vigreux column [a glass tube in which have been made a number of pairs of indentations which almost touch each other and which slope slightly downwards. The pairs of indentations are arranged to form a spiral of glass inside the tube].

For heating baths, see Table 3. For distillation apparatus on a macro, semi-micro or micro scale see Aldrich and other glassware catalogues. Alternatively, visit some useful websites for suppliers of laboratory glassware, e.g. <[www.wheatonsci.com](http://www.wheatonsci.com)>; <[www.sigmaldrich.com](http://www.sigmaldrich.com)> and <[www.kimble-kontes.com](http://www.kimble-kontes.com)>.

**Types of columns and packings.** A slow distillation rate is necessary to ensure that equilibrium conditions operate and also that the vapour does not become superheated so that the temperature rises above the boiling point. Efficiency is improved if the column is heat insulated (either by vacuum jacketing or by lagging) and, if necessary, heated to just below the boiling point of the most volatile component. Efficiency of separation also improves with increase in the heat of vaporisation of the liquids concerned (because fractionation depends on heat equilibration at multiple liquid-gas boundaries). Water and alcohols are more easily purified by distillation for this reason.

Columns used in distillation vary in their shapes and types of packing. Packed columns are intended to give efficient separation by maintaining a large surface of contact between liquid and vapour. Efficiency of separation is further increased by operation under conditions approaching total reflux, i.e. under a high reflux ratio. However, great care must be taken to avoid flooding of the column during distillation. The minimum number of theoretical plates for satisfactory separation of two liquids differing in boiling point by  $\delta t$  is approximately  $(273 + t)/3\delta t$ , where  $t$  is the average boiling point in °C.

Some of the commonly used *columns* are:

**Bruun column.** A type of all-glass bubble-cap column.

**Bubble-cap column.** A type of plate column in which inverted cups (bubble caps) deflect ascending vapour through reflux liquid lying on each plate. Excess liquid from any plate overflows to the plate lying below it and ultimately returns to the flask. (For further details, see Bruun & Faulconer *Ind Eng Chem (Anal Ed)* **9** 247 1937). Like most plate columns, it has a high through-put, but a relatively low number of theoretical plates for a given height.

**Dufton column.** A plain tube, into which fits closely (preferably ground to fit) a solid glass spiral wound round a central rod. It tends to choke at temperatures above 100° unless it is lagged (Dufton *J Soc Chem Ind (London)* **38** 45T 1919).

**Hempel column.** A plain tube (fitted near the top with a side arm) which is almost filled with a suitable packing, which may be of rings or helices.

**Oldershaw column.** An all-glass perforated-plate column. The plates are sealed into a tube, each plate being equipped with a baffle to direct the flow of reflux liquid, and a raised outlet which maintains a definite liquid level on the plate and also serves as a drain on to the next lower plate [see Oldershaw *Ind Eng Chem (Anal Ed)* **11** 265 1941].

**Podbielniak column.** A plain tube containing "Heli-Grid" Nichrome or Inconel wire packing. This packing provides a number of passage-ways for the reflux liquid, while the capillary spaces ensure very even spreading of the liquid, so that there is a very large area of contact between liquid and vapour while, at the same time, channelling and flooding are minimised. A column 1m high has been stated to have an efficiency of 200-400 theoretical plates (for further details, see Podbielniak *Ind Eng Chem (Anal Ed)* **13** 639 1941; Mitchell & O'Gorman *Anal Chem* **20** 315 1948).

**Stedman column.** A plain tube containing a series of wire-gauze discs stamped into flat, truncated cones and welded together, alternatively base-to-base and edge-to-edge, with a flat disc across each base. Each cone has a hole, alternately arranged, near its base, vapour and liquid being brought into intimate contact on the gauze surfaces (Stedman *Can J Research B* **15** 383 1937).

**Todd column.** A column (which may be a Dufton type, fitted with a Monel metal rod and spiral, or a Hempel type, fitted with glass helices) which is surrounded by an open heating jacket so that the temperature can be adjusted to be close to the distillation temperature (Todd *Ind Eng Chem (Anal Ed)* **17** 175 1945).

**Vigreux column.** A glass tube in which have been made a number of pairs of indentations which almost touch each other and which slope slightly downwards. The pairs of indentations are arranged to form a spiral of glass inside the tube.

**Widmer column.** A Dufton column, modified by enclosing within two concentric tubes the portion containing the glass spiral. Vapour passes up the outer tube and down the inner tube before entering the centre portion. Thus flooding of the column, especially at high temperatures, is greatly reduced (Widmer *Helv Chim Acta* **7** 59 1924).

The packing of a column greatly increases the surface of liquid films in contact with the vapour phase, thereby increasing the efficiency of the column, but reducing its capacity (the quantities of vapour and liquid able to flow in opposite directions in a column without causing flooding). Material for packing should be of uniform size, symmetrical shape, and have a unit diameter less than one-eighth that of the column. (Rectification efficiency increases sharply as the size of the packing is reduced but so, also, does the hold-up in the column.) It should also be capable of uniform, reproducible packing.

The usual *packings* are:

(a) **Rings.** These may be hollow glass or porcelain (Raschig rings), of stainless steel gauze (Dixon rings), or hollow rings with a central partition (Lessing rings) which may be of porcelain, aluminium, copper or nickel.

(b) **Helices.** These may be of metal or glass (Fenske rings), the latter being used where resistance to chemical attack is important (e.g. in distilling acids, organic halides, some sulphur compounds, and phenols). Metal single-turn helices are available in aluminium, nickel or stainless steel. Glass helices are less efficient, because they cannot be tamped to ensure uniform packing.

(c) **Balls.** These are usually glass.

(d) **Wire packing.** For use of "Heli-Grid" and "Heli-Pak" packings, see references given for Podbielniak column. For Stedman packing, see entry under Stedman column.

#### *Types of condensers:*

**Air condenser.** A glass tube such as the inner part of a Liebig condenser. Used for liquids with boiling points above 90°. Can be of any length.

**Allihn condenser.** The inner tube of a Liebig condenser is modified by having a series of bulbs to increase the condensing surface. Further modifications of the bubble shapes give the Julian and Allihn-Kronbitter condensers.

**Bailey-Walker condenser.** A type of all-metal condenser fitting into the neck of extraction apparatus and being supported by the rim. Used for high-boiling liquids.

**Coil condenser.** An open tube, into which is sealed a glass coil or spiral through which water circulates. The tube is sometimes also surrounded by an outer cooling jacket.

**Double surface condenser.** A tube in which the vapour is condensed between an outer and inner water-cooled jacket after impinging on the latter. Very useful for liquids boiling below 40°.

**Friedrichs condenser.** A "cold-finger" type of condenser sealed into a glass jacket open at the bottom and near the top. The cold finger is formed into glass screw threads.

**Graham condenser.** A type of coil condenser.

**Hopkins condenser.** A cold-finger type of condenser resembling that of Friedrichs.

**Liebig condenser.** An inner glass tube surrounded by a glass jacket through which water is circulated.

**Othmer condenser.** A large-capacity condenser which has two coils of relatively large bore glass tubing inside it, through which the water flows. The two coils join at their top and bottom.

**West condenser.** A Liebig condenser with a light-walled inner tube and a heavy-walled outer tube, with only a narrow space between them.

**Wiley condenser.** A condenser resembling the Bailey-Walker type.

**Vacuum distillation.** This expression is commonly used to denote a distillation under reduced pressure lower than that of the normal atmosphere. As the boiling point of a substance depends on the pressure, it is often possible to distil materials at a temperature low enough to avoid partial or complete decomposition by lowering the pressure, even if they are unstable when boiled at atmospheric pressure.

Sensitive or high-boiling liquids should invariably be distilled or fractionally distilled under reduced pressure. The apparatus is essentially as described for distillation except that ground joints connecting the different parts of the apparatus should be air tight by using grease, or better Teflon sleeves. For low, moderately high, and very high temperatures Apiezon L, M and T greases, respectively, are very satisfactory. Alternatively, it is often preferable to avoid grease and to use thin Teflon sleeves in the joints. The distilling flask, must be supplied with a capillary bleed (which allows a fine stream of air, nitrogen or argon into the flask), and the receiver should be of the fraction collector type. When distilling under vacuum it is very important to place a loose packing of glass wool above the liquid to buffer sudden boiling of the liquid. The flask should be not more than two-thirds full of liquid. The vacuum must have attained a steady state, i.e. the liquid has been completely degassed, before the heat source is applied, and the temperature of the heat source must be raised *very slowly* until boiling is achieved.

If the pump is a filter pump off a high-pressure water supply, its performance will be limited by the temperature of the water because the vapour pressure of water at 10°, 15°, 20° and 25° is 9.2, 12.8, 17.5 and 23.8 mmHg, respectively. The pressure can be measured with an ordinary manometer. For vacuums in the range 10<sup>-2</sup> mmHg to 10 mmHg, rotary mechanical pumps (oil pumps) are used and the pressure can be measured with a Vacustat McLeod-type gauge. If still higher vacuums are required, for example for high vacuum sublimations, a mercury diffusion pump is suitable. Such a pump can provide a vacuum up to 10<sup>-6</sup> mmHg. For better efficiencies, the diffusion pump can be backed up by a mechanical pump. In all cases, the mercury pump is connected to the distillation apparatus through several traps to remove mercury vapours. These traps may operate by chemical action, for example the use of sodium hydroxide pellets to react with acids, or by condensation, in which case empty tubes cooled in solid carbon dioxide-ethanol or liquid nitrogen (contained in wide-mouthed Dewar flasks) are used.

Special oil or mercury traps are available commercially, and a liquid-nitrogen (**b** -209.9°C) trap is the most satisfactory one to use between these and the apparatus. It has an advantage over liquid air or oxygen in that it is non-explosive if it becomes contaminated with organic matter. Air should not be sucked through the apparatus before starting a distillation because this will cause liquid oxygen (**b** -183°C) to condense in the liquid nitrogen trap, and this is potentially explosive (especially in mixtures with organic materials). Due to the potential lethal consequences of liquid oxygen/organic material mixtures, care must be exercised when handling liquid nitrogen. Hence, it is advisable to degas the system for a short period before the trap is immersed into the liquid nitrogen (which is kept in a Dewar flask).

**Spinning-band distillation.** Factors which limit the performance of distillation columns include the tendency to flood (which occurs when the returning liquid blocks the pathway taken by the vapour through the column) and the increased hold-up (which decreases the attainable efficiency) in the column that should, theoretically, be highly efficient. To overcome these difficulties, especially for distillation under high vacuum of heat sensitive or high-boiling highly viscous fluids, spinning band columns are commercially available. In such units, the distillation columns contain a rapidly rotating, motor-driven, spiral band, which may be of polymer-coated metal, stainless steel or platinum. The rapid rotation of the band in contact with the walls of the still gives intimate mixing of descending liquid with ascending vapour while the screw-like motion

of the band drives the liquid towards the still-pot, helping to reduce hold-up. There is very little pressure drop in such a system, and very high throughputs are possible, with high efficiency. For example, a 765-mm long 10-mm diameter commercial spinning-band column is reported to have an efficiency of 28 plates and a pressure drop of 0.2 mmHg for a throughput of 330ml/hour. The columns may be either vacuum jacketed or heated externally. The stills can be operated down to  $10^{-5}$  mmHg. The principle, which was first used commercially in the Podbielniak Centrifugal Superfractionator, has also been embodied in descending-film molecular distillation apparatus.

**Steam distillation.** When two immiscible liquids distil, the sum of their (independent) partial pressures is equal to the atmospheric pressure. Hence in steam distillation, the distillate has the composition

$$\frac{\text{Moles of substance}}{\text{Moles of water}} = \frac{P_{\text{substance}}}{P_{\text{water}}} = \frac{760 - P_{\text{water}}}{P_{\text{water}}}$$

where the  $P$ 's are vapour pressures (in mmHg) in the boiling mixture.

The customary technique consists of heating the substance and water in a flask (to boiling), usually with the passage of steam, followed by condensation and separation of the aqueous and non-aqueous phases in the distillate. Its advantages are those of selectivity (because only some water-insoluble substances, such as naphthalene, nitrobenzene, phenol and aniline are volatile in steam) and of ability to distil certain high-boiling substances well below their boiling point. It also facilitates the recovery of a non-steam-volatile solid at a relatively low temperature from a high-boiling solvent such as nitrobenzene. The efficiency of steam distillation is increased if superheated steam is used (because the vapour pressure of the organic component is increased relative to water). In this case the flask containing the material is heated (without water) in an oil bath and the steam passing through it is superheated by prior passage through a suitable heating device (such as a copper coil heated electrically or an oil bath).

**Azeotropic distillation.** In some cases two or more liquids form constant-boiling mixtures, or azeotropes. Azeotropic mixtures are most likely to be found with components which readily form hydrogen bonds or are otherwise highly associated, especially when the components are dissimilar, for example an alcohol and an aromatic hydrocarbon, but have similar boiling points.

Examples where the boiling point of the distillate is a minimum (less than either pure component) include:

**Water** with ethanol, *n*-propanol and isopropanol, *tert*-butanol, propionic acid, butyric acid, pyridine,  
**methanol** with methyl iodide, methyl acetate, chloroform,  
**ethanol** with ethyl iodide, ethyl acetate, chloroform, benzene, toluene, methyl ethyl ketone,  
**benzene** with cyclohexane,  
**acetic acid** with toluene.

Although less common, azeotropic mixtures are known which have higher boiling points than their components. These include water with most of the mineral acids (hydrofluoric, hydrochloric, hydrobromic, perchloric, nitric and sulfuric) and formic acid. Other examples are acetic acid-pyridine, acetone-chloroform, aniline-phenol, and chloroform-methyl acetate.

The following azeotropes are important commercially for drying ethanol:

ethanol 95.5% (by weight) - water 4.5%	<b>b</b> 78.1°
ethanol 32.4% - benzene 67.6%	<b>b</b> 68.2°
ethanol 18.5% - benzene 74.1% - water 7.4%	<b>b</b> 64.9°

Materials are sometimes added to form an azeotropic mixture with the substance to be purified. Because the azeotrope boils at a different temperature, this facilitates separation from substances distilling in the same range as the pure material. (Conversely, the impurity might form the azeotrope and be removed in this way.) This method is often convenient, especially where the impurities are isomers or are otherwise closely related to the desired substance. Formation of low-boiling azeotropes also facilitates distillation.

One or more of the following methods can generally be used for separating the components of an azeotropic mixture:

1. By using a chemical method to remove most of one species prior to distillation. (For example, water can be removed by suitable drying agents; aromatic and unsaturated hydrocarbons can be removed by sulfonation).
2. By redistillation with an additional substance which can form a ternary azeotropic mixture (as in the ethanol-water-benzene example given above).
3. By selective adsorption of one of the components. (For example, of water on to silica gel or molecular sieves, or of unsaturated hydrocarbons onto alumina).
4. By fractional crystallisation of the mixture, either by direct freezing or by dissolving in a suitable solvent.

**Kügelrohr distillation.** The Aldrich Kügelrohr Distillation Apparatus (see Aldrich-Sigma Labware catalogue) is made up of small glass bulbs (*ca* 4–5cm diameter) which are joined together *via* Quickfit joints at each pole of the bulbs. The liquid (or low melting solid) to be purified is placed in the first bulb of a series of bulbs joined end to end, and the system can be evacuated. The first bulb is heated in a furnace (e.g. Büchi Kügelrohr micro distillation oven from Sigma-Aldrich Labware catalogue) at a high temperature whereby most of the material distils into the second bulb (which is outside of the furnace). The second bulb is then moved into the furnace and the furnace temperature is reduced by *ca* 5° whereby the liquid in the second bulb distils into the third bulb (at this stage the first bulb is now out at the back of the furnace, and the third and subsequent bulbs are outside the front of the furnace). The furnace temperature is lowered by a further *ca* 5°, and the third bulb is moved into the furnace. The lower boiling material will distil into the fourth bulb. The process is continued until no more material distils into the subsequent bulb. The vacuum (if applied) and the furnace are removed, the bulbs are separated and the various fractions of distillates are collected from the individual bulbs. For volatile liquids, it may be necessary to cool the receiving bulb with solid CO<sub>2</sub> held in a suitable container (a Kügelrohr distillation apparatus with an integrated cooling system is available). This procedure is used for preliminary purification and the distillates are then redistilled or recrystallised.

**Isopiestic or isothermal distillation.** This technique can be useful for the preparation of metal-free solutions of volatile acids and bases for use in trace metal studies. The procedure involves placing two beakers, one of distilled water and the other of a solution of the material to be purified, in a desiccator. The desiccator is sealed and left to stand at room temperature for several days. The volatile components distribute themselves between the two beakers whereas the non-volatile contaminants remain in the original beaker. This technique has afforded metal-free pure solutions of ammonia, hydrochloric acid and hydrogen fluoride.

## RECRYSTALLISATION

### Techniques

The most commonly used procedure for the purification of a solid material by recrystallisation from a solution involves the following steps:

- (a) The impure material is dissolved in a suitable solvent, by shaking or vigorous stirring, at or near the boiling point, to form a near-saturated solution.
- (b) The hot solution is filtered to remove any insoluble particles. To prevent crystallisation during this filtration, a heated filter funnel can be used, or the solution can be diluted with more of the solvent.
- (c) The solution is then allowed to cool so that the dissolved substance crystallises out.
- (d) The crystals are separated from the mother liquor, either by centrifuging or by filtering, under suction, through a sintered glass, a Hirsch or a Büchner, funnel. Usually, centrifugation is preferred because of the greater ease and efficiency of separating crystals and mother liquor, and also because of the saving of time and effort, particularly when very small crystals are formed or when there is entrainment of solvent.
- (e) The crystals are washed free from mother liquor with a little fresh cold solvent, then dried.

If the solution contains extraneous coloured material likely to contaminate the crystals, this can often be removed by adding some activated charcoal (decolorising carbon) to the hot, but not boiling, solution which is then shaken frequently for several minutes before being filtered. (The large active surface of the carbon makes it a good adsorbent for this purpose.) In general, the cooling and crystallisation steps should be rapid so

as to give small crystals which occlude less of the mother liquor. This is usually satisfactory with inorganic material, so that commonly the filtrate is cooled in an ice-water bath while being vigorously stirred. In many cases, however, organic molecules crystallise much more slowly, so that the filtrate must be set aside to cool to room temperature or left in the refrigerator. It is often desirable to subject material that is very impure to preliminary purification, such as steam distillation, Soxhlet extraction, or sublimation, before recrystallising it. A greater degree of purity is also to be expected if the crystallisation process is repeated several times, especially if different solvents are used. The advantage of several crystallisations from different solvents lies in the fact that the material sought, and its impurities, are unlikely to have similar solubilities as solvents and temperatures are varied.

For the final separation of solid material, sintered-glass discs are preferable to filter paper. Sintered glass is unaffected by strongly acidic solutions or by oxidising agents. Also, with filter paper, cellulose fibres are likely to become included in the sample. The sintered-glass discs or funnels can be readily cleaned by washing in freshly prepared *chromic acid cleaning mixture*. This mixture is made by adding 100ml of concentrated sulfuric acid slowly with stirring to a solution of 5g of sodium dichromate (CARE: cancer suspect) in 5ml of water. (The mixture warms to about 70°, and sulfuric acid becomes hot when water is added to it; see p 3).

For materials with very low melting points it is sometimes convenient to use dilute solutions in acetone, methanol, pentane, diethyl ether or CHCl<sub>3</sub>/CCl<sub>4</sub>. The solutions are cooled to -78° in a dry-ice/acetone bath, to give a slurry which is filtered off through a precooled Büchner funnel. Experimental details, as applied to the purification of nitromethane, are given by Parrett and Sun [*J Chem Educ* 54 448 1977].

Where substances vary little in solubility with temperature, *isothermal crystallisation* may sometimes be employed. This usually takes the form of a partial evaporation of a saturated solution at room temperature by leaving it under reduced pressure in a desiccator.

However, in rare cases, crystallisation is not a satisfactory method of purification, especially if the impurity forms crystals that are isomorphous with the material being purified. In fact, the impurity content may even be greater in such recrystallised material. For this reason, it still remains necessary to test for impurities and to remove or adequately lower their concentrations by suitable chemical manipulation prior to recrystallisation.

**Filtration.** Filtration removes particulate impurities rapidly from liquids and is also used to collect insoluble or crystalline solids which separate or crystallise from solution. The usual technique is to pass the solution, cold or hot, through a fluted filter paper in a conical glass funnel.

If a solution is hot and needs to be filtered rapidly, a Büchner funnel and flask are used and filtration is performed under a slight vacuum (water pump), the filter medium being a circular cellulose filter paper wet with solvent. If filtration is slow, even under high vacuum, a pile of about twenty filter papers, wet as before, are placed in the Büchner funnel and, as the flow of solution slows down, the upper layers of the filter paper are progressively removed. Alternatively, a filter aid, e.g. Celite, Florisil or Hyflo-supercel, is placed on top of a filter paper in the funnel. When the flow of the solution (under suction) slows down, the upper surface of the filter aid is scratched gently. Filter papers with various pore sizes are available covering a range of filtration rates. Hardened filter papers are slow filtering, but they can withstand acidic and alkaline solutions without appreciable hydrolysis of the cellulose (see Table 4). When using strong acids it is preferable to use glass micro fibre filters, which are commercially available (see Tables 4 and 5).

Freeing a solution from extremely small particles [e.g. for optical rotatory dispersion (ORD) or circular dichroism (CD) measurements] requires filters with very small pore size. Commercially available (Millipore, Gelman, Nucleopore) filters other than cellulose or glass include nylon, Teflon, and polyvinyl chloride, and the pore diameter may be as small as 0.01 micron (see Table 9). Special containers are used to hold the filters, through which the solution is pressed by applying pressure, e.g. from a syringe. Some of these filters can be used to clear strong sulfuric acid solutions.

As an alternative to the Büchner funnel for collecting crystalline solids, a funnel with a sintered glass-plate under suction may be used. Sintered-glass funnels with various porosities are commercially available and can be easily cleaned with warm chromic or nitric acid (see above).

When the solid particles are too fine to be collected on a filter funnel because filtration is extremely slow, separation by **centrifugation** should be used. Bench-type centrifuges are most convenient for this purpose. The solid is placed in the centrifuge tube, the tubes containing the solutions on opposite sides of the rotor should be balanced accurately (at least within 0.05 to 0.1g), and the solutions are spun at maximum speed for as long as it takes to settle the solid (usually *ca* 3–5 minutes). The solid is washed (by shaking) with cold solvent by centrifugation, and finally twice with a pure volatile solvent in which the solid is insoluble, also by centrifugation.

After decanting the supernatant, the residue is dried in a vacuum, at elevated temperatures if necessary. In order to avoid "spitting" and contamination with dust while the solid in the centrifuge tube is dried, the mouth of the tube is covered with aluminium foil and held fast with a tight rubber band near the lip. The flat surface of the aluminium foil is then perforated in several places with a pin, and the tube and contents are dried in a vacuum desiccator over a desiccant.

### Solvents

**Choice of solvents.** The best solvents for recrystallisation have the following properties:

- (a) The material is much more soluble at higher temperatures than it is at room temperature or below.
- (b) Well-formed (but not large) crystals are produced.
- (c) Impurities are either very soluble or only sparingly soluble.
- (d) The solvent must be readily removed from the purified material.
- (e) There must be no reaction between the solvent and the substance being purified.
- (f) The solvent must not be inconveniently volatile or too highly flammable. (These are reasons why diethyl ether and carbon disulfide are not commonly used in this way.)

The following generalisations provide a rough guide to the selection of a suitable solvent:

- (a) Substances usually dissolve best in solvents to which they are most closely related in chemical and physical characteristics. Thus, hydroxylic compounds are likely to be most soluble in water, methanol, ethanol, acetic acid or acetone. Similarly, petroleum ether might be used with water-insoluble substances. However, if the resemblance is too close, solubilities may become excessive.
- (b) Higher members of homologous series approximate more and more closely to their parent hydrocarbon.
- (c) Polar substances are more soluble in polar than in non-polar solvents.

Although Chapters 4, 5 and 7 provide details of the solvents used for recrystallising a large portion of commercially available laboratory chemicals, they cannot hope to be exhaustive, nor need they necessarily be the best choice, but they are the solvents reported in the literature. In other cases where it is desirable to use this process, it is necessary to establish whether a given solvent is suitable. This is usually done by taking only a small amount of material in a small test-tube and adding enough solvent to cover it. If it dissolves readily in the cold or on gentle warming, the solvent is unsuitable. Conversely, if it remains insoluble when the solvent is heated to boiling (adding more solvent if necessary), the solvent is again unsuitable. If the material dissolves in the hot solvent but does not crystallise readily within several minutes of cooling in an ice-salt mixture, another solvent should be tried.

### Water

The properties and purification of water are described in the "Inorganic Compounds" section of Chapter 5. Fluka (Riedel-de Haën) supply purified water prepared specifically for a variety of uses, e.g. LC-MS, HPLC, gradient elution, for cell biology which is freed from enterotoxins by ultrafiltration and autoclaving, for organic and for inorganic trace analysis, for residue analysis and other analytical purposes. Some of these have been prepared by reverse osmosis, or ultrafiltration, under clean room conditions and filtered through 0.2  $\mu\text{m}$  membranes into bottles of high purity glass under inert gas. They have a limited shelf life once opened most probably because  $\text{O}_2$  from the air dissolves readily in the water. The solubility of  $\text{O}_2$  in 100 ml of water is ~1.02 ml (0.455 mM) at 0°, 0.68 ml (0.282 mM) at 20°, 0.63 ml (0.258 mM) at 25°, 0.63 ml (0.237 mM) at 30°, and 0.12 ml (0.033 mM) at 100°, all at ~760 mmHg in equilibrium with air (see Tables 20–23). This is in comparison with the concentration of  $\text{O}_2$  of 0.23 mM in 0.1M Tris HCl buffer at pH 7.2 and 25° in equilibrium with air at 760 mmHg. Routinely, water is best purified by redistilling it twice in an all glass apparatus, storing it under  $\text{N}_2$  or He in stoppered glass containers and, if necessary, preferably subjected to ultrafiltration through a single or multistage 0.2  $\mu\text{m}$  membrane system or reverse osmosis (visit <[www.millipore.com](http://www.millipore.com)>). If oxygen-free water is required,  $\text{N}_2$  or argon should be bubbled through a sintered glass frit in the highly purified water for 2–3 hours, and stoppered immediately. It is best to use a glass container from which the water can be withdrawn without it coming into contact with air. Note that boiling and distilling water, and condensing it in an inert atmosphere should de-gas it.

**Petroleum ethers** are commercially available fractions of refined petroleum and are sold in fractions of about 20° boiling ranges. This ensures that little of the hydrocarbon ingredients boiling below the range is lost during standing or boiling when recrystallising a substance. Petroleum ethers with boiling ranges (at 760 mm pressure) of 35–60°, 40–60°, 60–80°, 80–100°, and 100–120° are generally free from unsaturated and aromatic hydrocarbons. The lowest boiling petroleum ether commercially available has **b** 30–40°/760 mm and is mostly *n*-pentane. The purer spectroscopic grades are almost completely free from olefinic and aromatic hydrocarbons. **Petroleum spirit** (which is sometimes used synonymously with petroleum ether or light petroleum) is usually less refined petroleum, and **ligroin** is used for fractions boiling above 100°. The lower boiling fractions consist of mixtures of *n*-pentane (**b** 36°), *n*-hexane (**b** 68.5°) and *n*-heptane (**b** 98°), and some of their isomers in varying proportions. For purification see petroleum ether **b** 35–60° in “Aliphatic Compounds”, chapter 4, which is typical.

Solvents commonly used for recrystallisation, and their boiling points, are given in Table 6. For comments on the toxicity and use of **benzene** see the “Introduction” pages of Chapters 4, 5 and 6.

**Mixed Solvents.** Where a substance is too soluble in one solvent and too insoluble in another, for either to be used for recrystallisation, it is often possible (provided the solvents are miscible) to use them as a mixed solvent system. (In general, however, it is preferable to use a single solvent if this is practicable.) Table 7 contains many of the common pairs of miscible solvents.

Several procedures with mixed solvents have been used successfully for crystallisation. These include the following:

(a) The material is dissolved in the solvent in which it is more soluble at room temperature, then the second solvent (heated to near boiling) is added cautiously to the cold solution until a slight turbidity persists or crystallisation begins. The turbidity is cleared by warming or by adding several drops of the first solvent, and the clear solution is allowed to cool slowly for crystallisation to occur. The supernatant is decanted off carefully (do not disturb the crystals unduly) and more of the second solvent is added to the clear decanted supernatant until turbidity begins again, and is set aside for further crystals to form. The procedure is repeated until no more crystals separate.

(b) A variation of the procedure in (a) is simply to precipitate the material in a microcrystalline form from solution in one solvent at room temperature, by adding a little more of the second solvent also at room temperature, filtering off the crystals, adding a little more of the second solvent and repeating the process. This ensures, at least in the first or last precipitation, a material which contains as little as possible of the impurities, which may also be precipitated in this way. With inorganic salts or metal salts of organic acids, the first solvent is commonly water, and the second solvent is alcohol or acetone. With salts of organic bases and inorganic acids, e.g. hydrochloride, or salts of organic acids and organic bases, the first solvent is usually an alcohol or acetone, in which the salt is very soluble and the second solvent is dry diethyl ether.

(c) A very concentrated solution of the compound in the first solvent in one beaker, and a second beaker containing the second solvent in which the compound is insoluble are placed in a desiccator. As the vapours of the two solvents equilibrate in the desiccator, and crystals separate in the first beaker that contains the compound.

(d) This procedure is best carried out in a cold room (at *ca* 4°). A strong solution of the solid in the solvent in which it is very soluble is *layered* carefully with the second solvent. As the second solvent diffuses and dissolves into the solution, crystals begin to form at the ‘interface’. When separation of crystals is complete and the solvent mixture is homogeneous, another layer of the second solvent is applied and the process is repeated.

**Seeding** is well known to initiate the crystallisation process. A good way to procure seed crystals is to dissolve the crystals in the minimum amount of solvent, place the solution in a watch glass, then blow a fine stream of dry N<sub>2</sub> or argon gently over the surface of the solution until seed crystals are formed. Alternatively, the inert gas is allowed to evaporate all the solvent, and the residual crystals or fine powder are used for seeding. The seeds are applied in the above procedures at the appropriate time, e.g. when first turbidity appears, or placed onto the tip of a glass rod which is then rubbed against the sides of the container of the solution until crystallisation begins.

### Recrystallisation from the melt

A crystalline solid melts when its temperature is raised sufficiently for the thermal agitation of its molecules or ions to overcome the restraints imposed by the crystal lattice. Usually, impurities weaken crystal structures, and hence lower the melting points of solids (or the freezing points of liquids). If an impure material is melted and cooled slowly (with the addition, if necessary, of a trace of solid material near the freezing point to avoid supercooling), the first crystals that form will usually contain less of the impurity, so that fractional solidification by partial freezing can be used as a purification process for solids with melting points lying in a convenient temperature range (or for more readily frozen liquids). Some examples of cooling baths that are

useful in recrystallisation are summarised in Table 8. In some cases, impurities form higher melting eutectics with substances to be purified, so that the first material to solidify is less pure than the melt. For this reason, it is often desirable to discard the first crystals and also the final portions of the melt. Substances having similar boiling points often differ much more in melting points, so that fractional solidification can offer real advantages, especially where ultrapurity is sought. For further information on this method of recrystallisation, consult the earlier editions of this book as well as references by Schwab and Wichers (*J Res Nat Bur Stand* **25** 747 1940). This method works best if the material is already nearly pure, and hence tends to be a final purification step.

**Zone refining.** Zone refining (or zone melting) is a particular development for fractional solidification and is applicable to all crystalline substances that show differences in the concentrations of impurities in liquid and solid states at solidification. The apparatus used in this technique consists essentially of a device in which the crystalline solid to be purified is placed in a glass tube (set vertically) which is made mechanically to move slowly upwards while it passes through a fixed coil (one or two turns) of heated wire. A narrow zone of molten crystals is formed when the tube is close to the heated coil. As the zone moves away from the coil the liquid crystallises, and a fresh molten zone is formed below it at the coil position. The machine can be set to recycle repeatedly. At its advancing side, the zone has a melting interface with the impure material whereas on the upper surface of the zone there is a constantly growing face of higher-melting, resolidified purer material. This leads to a progressive increase in impurity in the liquid phase which, at the end of the run, is discarded from the bottom of the tube. Also, because of the progressive increase in impurity in the liquid phase, the resolidified material contains correspondingly less of the impurities. For this reason, it is usually necessary to make several zone-melting runs before a sample is satisfactorily purified. This is also why the method works most successfully if the material is already fairly pure. In all these operations the zone must travel slowly enough to enable impurities to diffuse or be convected away from the area where resolidification is occurring. The technique finds commercial application in the production of metals of extremely high purity (tubes other than glass are used in these cases, and impurities are reduced down to  $10^9$  ppm), in purifying refractory oxides, and in purifying organic compounds, using commercially available equipment. Criteria for indicating that definite purification is achieved include elevation of melting point, removal of colour, fluorescence or smell, and a lowering of electrical conductivity. Difficulties likely to be found with organic compounds, especially those of low melting points and low rates of crystallisation, are supercooling and, because of surface tension and contraction, the tendency of the molten zone to seep back into the recrystallised areas. The method is likely to be useful in cases where fractional distillation is not practicable, either because of unfavourable vapour pressures or ease of decomposition, or where super-pure materials are required. The method has been used for the latter purpose for purifying anthracene, benzoic acid, chrysene, morphine, 1,8-naphthyridine and pyrene to name a few. [See E.F.G. Herington, *Zone Melting of Organic Compounds*, Wiley & Sons, NY, 1963; W. Pfann, *Zone Melting*, 2nd edn, Wiley, NY, 1966; H. Schildknecht, *Zonenschmelzen*, Verlag Chemie, Weinheim, 1964; W.R. Wilcox, R.Friedenberg et al. *Chem Rev* **64** 187 1964; M. Zief and W.R. Wilcox (Eds), *Fractional Solidification*, Vol I, M Dekker Inc. NY, 1967.]

## SUBLIMATION

Sublimation differs from ordinary distillation because the vapour condenses to a solid instead of a liquid. Usually, the pressure in the heated system is diminished by pumping, and the vapour is condensed (after travelling a relatively short distance) onto a cold finger or some other cooled surface. This technique, which is applicable to many organic solids, can also be used with inorganic solids such as aluminium chloride, ammonium chloride, arsenious oxide and iodine to name a few. In some cases, passage of a stream of inert gas over the heated substance secures adequate vaporisation and reduces oxidation. This procedure has the added advantage of removing occluded solvent used for recrystallising the solid.

## CHROMATOGRAPHY

Chromatography is often used with advantage for the purification of small, and large, amounts of complex organic mixtures. Chromatography techniques all rely on the differential distribution of the various components in the solution, between the mobile phase and the stationary phase. The mobile phase can either be a gas or a liquid, whereas the stationary phase can either be a solid or a non-volatile liquid adsorbed on a solid surface.

The major chromatographic techniques can also be categorised according to the nature of the mobile phase used – vapour phase chromatography for when a gas is the mobile phase and liquid chromatography for when a liquid is the mobile phase.

The suppliers of chromatography equipment for every need are too numerous to list here but can be viewed on the internet under “Chromatography products”. Details and orders can be obtained from the respective websites listed at the end of the section on HPLC below.

### Vapour phase chromatography (GC or gas-liquid chromatography)

The mobile phase in vapour phase chromatography is a gas (e.g. hydrogen, helium, nitrogen or argon), and the stationary phase is a non-volatile liquid impregnated onto a porous material. The mixture to be purified is injected into a heated inlet whereby it is vaporised and taken into the column by the carrier gas. It is separated into its components by partition between the liquid on the porous support and the gas. For this reason vapour-phase chromatography is sometimes referred to as gas-liquid chromatography (g.l.c). Vapour phase chromatography is very useful for the resolution of a mixture of volatile compounds. This type of chromatography uses either packed or capillary columns. Packed columns have internal diameters of 3–5 mm with lengths of 2–6 metres. These columns can be packed with a range of materials including firebrick derived materials (chromasorb P, for separation of non-polar hydrocarbons) or diatomaceous earth (chromasorb W, for separation of more polar molecules such as acids, amines). Capillary columns have stationary phase bonded to the walls of long capillary tubes. The diameters of capillary columns are less than 0.5 mm, and the lengths of these columns can go up to 50 metres! These columns have much superior separating powers than the packed columns. Elution times for equivalent resolutions with packed columns can be up to ten times shorter. It is believed that almost any mixture of compounds can be separated using one of the four stationary phases, OV-101, SE-30, OV-17 and Carbowax-20M. Capillary columns for analysis in gas chromatography are now routinely used. An extensive range of packed and capillary columns is available from chromatographic specialists such as Supelco, Alltech, Hewlett-Packard, Phenomenex (for stainless steel capillary columns see <phenomenex.com>, etc. (see above and at the end of the section on HPLC below).

Table 9 shows some typical liquids used for stationary phases in gas chromatography.

Although gas chromatography is routinely used for the analysis of mixtures, this form of chromatography can also be used for separation/purification of substances. This is known as preparative GC. In preparative GC, suitably packed columns are used, and as substances emerge from the column, they are collected by condensing the vapour of these separated substances in suitable traps. The carrier gas blows the vapour through these traps; hence these traps have to be very efficient. Improved collection of the effluent vaporised fractions in preparative work is attained by strong cooling, increasing the surface of the traps by packing them with glass wool, and/or by applying an electrical potential which neutralises the charged vapour and causes it to condense.

When the gas chromatograph is attached to a mass spectrometer, a very powerful analytical tool (*gas chromatography-mass spectrometry*; **GC-MS**) is produced. Gas chromatography allows the separation of mixtures but does not allow the definitive identification of unknown substances, whereas mass spectrometry is good for the identification of individual compounds of the mixtures of compounds. This means that with GC-MS, both separation *and* identification of substances in mixtures can be achieved. The spectrometer can be connected to a computer that has a library from which the mass peaks can be compared and is a very powerful analytical tool. Because of the relatively small amounts of material required for mass spectrometry, a splitting system is inserted between the column and the mass spectrometer. This enables only a small fraction of the effluent to enter the spectrometer; the rest of the effluent is usually collected or vented to the air.

For more detail on apparatus and chromatographic columns see <http://www.sigmaldrich.com/analytical-chromatography/gas-chromatography.html> and websites at the end of the section on HPLC below.

### Liquid chromatography

In contrast to vapour phase chromatography, the mobile phase in liquid chromatography is a liquid. In general, there are four main types of liquid chromatography: *adsorption*, *partition*, *ion-chromatography*, and *gel filtration*.

**Adsorption chromatography** is based on the difference in the extent to which substances in solution are adsorbed onto a suitable surface. The main techniques in adsorption chromatography are TLC (thin layer chromatography), paper and column chromatography.

*Thin layer chromatography (TLC).* In thin layer chromatography, the mobile phase, i.e. the solvent, creeps up the stationary phase (the adsorbent) by capillary action. The adsorbent (e.g. silica, alumina, cellulose) is spread on a rectangular glass plate (or solid inert plastic sheet or aluminium foil). Some adsorbents (e.g. silica) are mixed with a setting material (e.g. CaSO<sub>4</sub>) by the manufacturers which causes the film to set hard on drying. The adsorbent can be activated by heating at 100–110° for a few hours. Other adsorbents (e.g. celluloses) adhere on glass plates without a setting agent. Thus some grades of adsorbents have prefixes; e.g. prefix G means that the adsorbent can cling to a glass plate and is used for TLC (e.g. silica gel GF254 is for TLC plates which have a dye that fluoresces under 254nm UV light). Those lacking this binder have the letter H after any coding and is suitable for column chromatography e.g. silica gel 60H. The materials to be purified or separated are spotted in a solvent close to the lower end of the plate and allowed to dry. The spots will need to be placed at such a distance so as to ensure that when the lower end of the plate is immersed in the solvent, the spots are a few mm above the eluting solvent. The plate is placed upright in a tank containing the eluting solvent. Elution is carried out in a closed tank to ensure equilibrium. Good separations can be achieved with square plates if a second elution is performed at right angles to the first using a second solvent system. For rapid work, plates of the size of microscopic slides or even smaller are used which can decrease the elution time and cost without loss of resolution. The advantage of plastic backed and aluminium foil backed plates is that the size of the plate can be made as required by cutting the sheet with scissors or a sharp guillotine. Visualisation of substances on TLC can be carried out using UV light if they are UV absorbing or fluorescing substances or by spraying or dipping the plate with a reagent that gives coloured products with the substance (e.g. iodine solution or vapour gives brown colours with amines), or with dilute sulfuric acid (organic compounds become coloured or black when the plates are heated at 100° if the plates are of alumina or silica, but not cellulose). (See Table 10 for some methods of visualisation.) Some alumina and silica powders are available with fluorescent materials in them, in which case the whole plate fluoresces under UV light. Non-fluorescing spots are thus clearly visible, and fluorescent spots invariably fluoresce with a different colour. The colour of the spots can be different under UV light at 254 nm and at 365 nm. Another useful way of showing up non-UV absorbing spots is to spray the plate with a 1–2% solution of Rhodamine 6G in acetone. Under UV light the dye fluoresces and reveals the non-fluorescing spots. For preparative work, if the material in the spot or fraction is soluble in ether or petroleum ether, the desired substance can be extracted from the adsorbent with these solvents which leave the water soluble dye behind.

TLC can be used as an analytical technique, or as a guide to establishing conditions for column chromatography or as a preparative technique in its own right.

The thickness of the adsorbent on the TLC plates could be between 0.2 mm to 2 mm or more. In preparative work, the thicker plates are used and hundreds of milligrams of mixtures can be purified conveniently and quickly. The spots or areas are easily scraped off the plates and the desired substances extracted from the adsorbent with the required solvent. For preparative TLC, non-destructive methods for visualising spots and fractions are required. As such, the use of UV light is very useful. If substances are not UV active, then a small section of the plate (usually the right or left edge of the plate) is sprayed with a visualising agent while the remainder of the plate is kept covered.

Thin layer chromatography has been used successfully with ion-exchange celluloses as stationary phases and various aqueous buffers as mobile phases. Also, gels (e.g. Sephadex G-50 to G-200 superfine) have been adsorbed on glass plates and are good for fractionating substances of high molecular weights (1500 to 250,000). With this technique, which is called *thin layer gel filtration (TLG)*, molecular weights of proteins can be determined when suitable markers of known molecular weights are run alongside (see Chapter 6).

Commercially available pre-coated plates with a variety of adsorbents are generally very good for quantitative work because they are of a standard quality. Plates of a standardised silica gel 60 (as medium porosity silica gel with a mean porosity of 6 mm) released by Merck have a specific surface of 500 m<sup>2</sup>/g and a specific pore volume of 0.75 ml/g. They are so efficient that they have been called *high performance thin layer chromatography (HPTLC)* plates (Ropphahn & Halpap *J Chromatogr* **112** 81 1975). In another variant of thin layer chromatography the adsorbent is coated with an oil as in gas chromatography thus producing *reverse-phase thin layer chromatography*. Reversed-phase TLC plates e.g. silica gel RP-18 are available from Fluka and Merck.

A very efficient form of chromatography makes use of a circular glass plate (rotor) coated with an adsorbent (silica, alumina or cellulose). As binding to a rotor is needed, the sorbents used may be of a special quality and/or binders are added to the sorbent mixtures. For example when silica gel is required as the adsorbent, silica gel 60 PF-254 with calcium sulfate (Merck catalog 7749) is used. The thickness of the adsorbent (1, 2 or 4mm) can vary depending on the amount of material to be separated. The apparatus used is called a **Chromatotron** (available from Harrison Research, USA). The glass plate is rotated by a motor, and the sample followed by the eluting solvent is allowed to drip onto a central position on the plate. As the plate rotates the solvent elutes the

mixture, centrifugally, while separating the components in the form of circular bands radiating from the central point. The separated bands are usually visualised conveniently by UV and as the bands approach the edge of the plate, the eluent is collected. The plate with the adsorbent can be re-used many times if care is employed in the usage, and hence this form of chromatography utilises less adsorbents as well as solvents.

Recipes and instructions for coating the rotors are available from the Harrison website <<http://www.harrisonresearch.com/chromatotron/>>. In addition, information on how to regenerate the sorbents and binders is also included.

*Paper chromatography.* This is the technique from which thin layer chromatography was developed. It uses cellulose paper (filter paper) instead of the TLC adsorbent and does not require a backing like the plastic sheet in TLC. It is used in the **ascending procedure** (like in TLC) whereby a sheet of paper is hung in a jar, and the materials to be separated are spotted (after dissolving in a suitable solvent and drying) near the bottom of the sheet which dips into the eluting solvent just below the spots. As the solvent rises up the paper the spots are separated according to their adsorption properties. A variety of solvents can be used, the sheet is then dried in air (fume cupboard), and can then be run again with the solvent running at right angles to the first run to give a two-dimensional separation. The spots can then be visualised as in TLC or can be cut out and analysed as required. A **descending procedure** had also been developed where the material to be separated is spotted near the top of the paper and the top end is made to dip into a tray containing the eluting solvent. The whole paper is placed in a glass jar, and the solvent then runs down the paper causing the materials in the spots to separate also according to their adsorption properties and to the eluting ability of the solvent. This technique is much cheaper than TLC and is still used (albeit with thicker cellulose paper) with considerable success for the separation of protein hydrolysates for sequencing analysis and/or protein identification. However, modern and more efficient technologies are available for analysing proteins and their hydrolysates although the equipment is expensive. (Whatman papers for chromatography and electrophoresis are available also from Sigma-Aldrich Labware.)

*Column Chromatography.* The substances to be purified are usually placed on the top of the column and the solvent is run down the column. Fractions are collected and checked for compounds using TLC (UV and/or other means of visualisation). The adsorbent for chromatography can be packed dry and solvents to be used for chromatography are used to equilibrate the adsorbent by flushing the column several times until equilibration is achieved. Alternatively, the column containing the adsorbent is packed wet (slurry method), and pressure is applied at the top of the column until the column is well packed (i.e. the adsorbent is settled).

*Graded Adsorbents and Solvents.* Some materials used in columns for adsorption chromatography are grouped in Table 11 in an approximate order of effectiveness. Other adsorbents sometimes used include barium carbonate, calcium sulfate, calcium phosphate, charcoal (usually mixed with Kieselguhr or other form of diatomaceous earth, for example, the filter aid Celite) and cellulose. The alumina can be prepared in several grades of activity (see below).

In most cases, adsorption takes place most readily from non-polar solvents such as petroleum ether and least readily from polar solvents such as alcohols, esters, and acetic acid. Common solvents, arranged in approximate order of increasing eluting ability are also given in Table 11. Eluting power roughly parallels the dielectric constants of solvents. The series also reflects the extent to which the solvent binds to the column material, thereby displacing the substances that are already adsorbed. This preference of alumina and silica gel for polar molecules explains, for example, the use of percolation through a column of silica gel for the following purposes—drying of ethylbenzene, removal of aromatics from 2,4-dimethylpentane and of ultraviolet absorbing substances from cyclohexane.

Mixed solvents are intermediate in strength, and so provide a finely graded series. In choosing a solvent for use as an eluent it is necessary to consider the solubility of the substance in it and the ease with which it can subsequently be removed.

*Preparation and Standardisation of Alumina.* The activity of alumina depends inversely on its water content, and a sample of poorly active material can be rendered more active by leaving for some time in a round bottomed flask heated up to about 200° in an oil bath or a heating mantle while a slow stream of a dry inert gas is passed through it. *Alternatively*, it is heated to red heat (380–400°) in an open vessel for 4–6 hours with occasional stirring and then cooled in a vacuum desiccator: this material is then of grade I activity. Conversely, alumina can be rendered less active by adding small amounts of water and thoroughly mixing for several hours. Addition of about 3% (w/w) of water converts grade I alumina to grade II.

Used alumina can be regenerated by repeated extraction, first with boiling methanol, then with boiling water, followed by drying and heating. The degree of activity of the material can be expressed conveniently in terms of the scale due to Brockmann and Schodder (*Chem Ber B* **74** 73 1941).

Alumina is normally slightly alkaline. A (less strongly adsorbing) neutral alumina can be prepared by making a slurry in water and adding 2M hydrochloric acid until the solution is acid to Congo red. The alumina is then filtered off, washed with distilled water until the wash water gives only a weak violet colour with Congo red paper, and dried.

Alumina used in TLC can be recovered by washing in ethanol for 48 hours with occasional stirring, to remove binder material and then washed with successive portions of ethyl acetate, acetone and finally with distilled water. Fine particles are removed by siphoning. The alumina is first suspended in 0.04M acetic acid, then in distilled water, siphoning off 30 minutes after each wash. The process is repeated 7–8 times. It is then dried and activated at 200° [Vogh & Thomson *Anal Chem* **53** 1365 1981].

#### *Preparation of other adsorbents*

*Silica gel* can be prepared from commercial water-glass by diluting it with water to a density of 1.19 and, while keeping it cooled to 5°, adding concentrated hydrochloric acid with stirring until the solution is acid to thymol blue. After standing for 3 hours, the precipitate is filtered off, washed on a Büchner funnel with distilled water, then suspended in 0.2M hydrochloric acid. The suspension is set aside for 2–3 days, with occasional stirring, then filtered, washed well with water and dried at 110°. It can be activated by heating up to about 200° as described for alumina.

Powdered commercial silica gel can be purified by suspending and standing overnight in concentrated hydrochloric acid (6 ml/g), decanting the supernatant and repeating with fresh acid until the latter remains colourless. After filtering with suction on a sintered-glass funnel, the residue is suspended in water and washed by decantation until free of chloride ions. It is then filtered, suspended in 95% ethanol, filtered again and washed on the filter with 95% ethanol. The process is repeated with anhydrous diethyl ether before the gel is heated for 24 hours at 100° and stored for another 24 hours in a vacuum desiccator over phosphorus pentoxide.

To buffer silica gel for flash chromatography (see later), 200g of silica is stirred in 1L of 0.2M NaH<sub>2</sub>PO<sub>4</sub> for 30 minutes. The slurry is then filtered with suction using a sintered glass funnel. The silica gel is then activated at 110°C for 16 hours. The pH of the resulting silica gel is ~4. Similar procedures can be utilised to buffer the pH of the silica gel at various pHs (up to pH ~8: pH higher than this causes degradation of silica) using appropriate phosphate buffers.

Commercial silica gel has also been purified by suspension of 200g in 2L of 0.04M ammonia, and stood for 5 minutes before siphoning off the supernatant. The procedure was repeated 3–4 times, before rinsing with distilled water and drying, and activating the silica gel in an oven at 110° [Vogh & Thomson, *Anal Chem* **53** 1345 1981].

Although silica gel is not routinely recycled after use (due to fear of contamination as well as the possibility of reduced activity), the costs of using new silica gel for purification may be prohibitive. In these cases, recycling may be achieved by stirring the used silica gel (1 kg) in a mixture of methanol and water (2L MeOH/4L water) for 30–40 minutes. The silica gel is filtered (as described above) and reactivated at 110°C for 16 hours.

*Diatomaceous earth* (Celite 535 or 545, Hyflo Super-cel, Dicalite, Kieselguhr) is purified before use by washing with 3M hydrochloric acid, then water, or it is made into a slurry with hot water, filtered at the pump and washed with water at 50° until the filtrate is no longer alkaline to litmus. Organic materials can be removed by repeated extraction at 50° with methanol or chloroform, followed by washing with methanol, filtering and drying at 90–100°.

*Charcoal* is generally satisfactorily activated by heating gently to red heat in a crucible or quartz beaker in a muffle furnace, finally allowing to cool under an inert atmosphere in a desiccator. Good commercial activated charcoal is made from wood, e.g. *Norit* (from Birch wood), *Darco* and *Nuchar*. If the cost is important, then the cheaper *animal charcoal* (bone charcoal) can be used. However, this charcoal contains calcium phosphate and other calcium salts and cannot be used with acidic materials. In this case the charcoal is boiled with dilute hydrochloric acid (1:1 by volume) for 2–3 hours, diluted with distilled water and filtered through a fine grade paper on a Büchner flask, washed with distilled water until the filtrate is almost neutral, and dried first in air, then in a vacuum, and activated as above. To improve the porosity, charcoal columns are usually prepared in admixture with diatomaceous earth.

*Cellulose* for chromatography is purified by sequential washing with chloroform, ethanol, water, ethanol, chloroform and acetone. More extensive purification uses aqueous ammonia, water, hydrochloric acid, water, acetone and diethyl ether, followed by drying in a vacuum. Trace metals can be removed from filter papers by washing for several hours with 0.1M oxalic, citric acid, or 0.1M EDTA solution, followed by repeated washing with distilled water.

Supelco supply a variety of “solvent desorption tubes”, which are cartridges that remove specific impurities (e.g. LpDNPH cartridges which contain a high purity silica adsorbent coated with 2,4-dinitrophenylhydrazine and remove carbonyl compounds; ozone scrubbers which eliminate ozone). Other cartridges such as the “ORBO charcoal” cartridges contain various beds such as *activated coconut charcoal*, *activated petroleum charcoal*, *HBr on petroleum charcoal* or *4-tert-butyl catechol on charcoal* and are used for specific or for general purposes. Other ORBO cartridges contain activated silica gel and coated silica gel, Florisil, Carboxen, Carbosieve and carbon-coated traps, as well as a variety of ORBO porous polymers, polyurethane, and glass fibre coated with 1-(2-pyridyl)piperazine (which is specific for sampling diisocyanates). They also supply filter cartridges for trapping aerosols and particulate forms of semivolatiles.

### Flash Chromatography (FC and HPFC)

A faster method of separating components of a mixture is *flash chromatography* (see Still et al. *J Org Chem* **43** 2923 1978). Flash chromatography has become an extremely useful and popular means of purification of small as well as large quantities of compounds. In flash chromatography the eluent flows through the column under a pressure of *ca* 1 to 4 atmospheres. The lower end of the chromatographic column has a relatively long taper closed with a tap. The upper end of the column is connected through a ball joint to a tap. Alternatively a specially designed chromatographic column with a solvent reservoir can also be used (for an example, see the Aldrich Chemical Catalog-glassware section). The tapered portion is plugged with cotton, or quartz, wool and *ca* 1 cm length of fine washed sand (the latter is optional). The adsorbent is then placed in the column as a dry powder or as a slurry in a solvent and allowed to fill to about one-third of the column. A fine grade of adsorbent is required in order to slow the flow rate at the higher pressure, e.g. Silica 60, 230 to 400 mesh with particle size 0.040-0.063mm (e.g. from Merck). The top of the adsorbent is layered with *ca* 1 cm length of fine washed sand. The mixture in the smallest volume of solvent is applied at the top of the column and allowed to flow into the adsorbent under gravity by opening the lower tap momentarily. The top of the column is filled with eluent, the upper tap is connected by a tube to a nitrogen supply from a cylinder, or to compressed air, and turned on to the desired pressure (monitor with a gauge). The lower tap is turned on and fractions are collected rapidly until the level of eluent has reached the top of the adsorbent (do not allow the column to run dry). If further elution is desired then both taps are turned off, the column is filled with more eluting solvent and the process repeated. The top of the column can be modified so that gradient elution can be performed. Alternatively, an apparatus for producing the gradient is connected to the upper tap by a long tube and placed high above the column in order to produce the required hydrostatic pressure. Much better resolution is obtained by dry loading the sample for purification rather than loading the sample as a solution. Flash chromatography is more efficient and gives higher resolution than conventional chromatography at atmospheric pressure and is completed in a relatively shorter time. A successful separation of components of a mixture by TLC using the same adsorbent is a good indication that flash chromatography will give the desired separation on a larger scale.

Very elaborate equipment is now available for FC and HPFC (high-performance flash chromatography), which may include a pump, facility for gradient elution, UV detection and fraction collection of effluent. A large variety of columns (disposable cartridges) with packings such as silicate, carbon, reverse phases for a wide range of applications are commercially available. In addition a plethora of cartridges are available for preliminary purification, prior to FC or HPFC, packed with adsorbents which can remove specific impurities, e.g. unwanted reaction products such as aldehydes or ketone which may be suspected by-products and/or starting materials See

Supelco online catalog <http://sigma-aldrich.dirxion.com/WebProject.asp?BookCode=chr09flx#>

Sigma Aldrich: <http://www.sigmaaldrich.com/analytical-chromatography/analytical-chromatography-catalog.html>

Biotage: *Synthesis and Purification Catalogue* and the *Analytical Sample Preparation Catalogue* contain details on available FC and HPFC equipment, accessories and consumables, as well as means of optimising purification, see <[www.biotage.com](http://www.biotage.com)>.

### Paired-ion Chromatography (PIC)

Mixtures containing ionic compounds (e.g. acids and/or bases), non-ionisable compounds, and zwitterions can be separated successfully by paired-ion chromatography (PIC). It utilises the “reverse-phase” technique (Eksberg & Schill *Anal Chem* **45** 2092 1973). The stationary phase is lipophilic, such as  $\mu$ -BONDAPAK C<sub>18</sub> or any other adsorbent that is compatible with water. The mobile phase is water or aqueous methanol containing the acidic or basic counter ion. Thus the mobile phase consists of dilute solutions of strong acids (e.g. 5mM

1-heptanesulfonic acid) or strong bases (e.g. 5 mM tetrabutylammonium phosphate) that are completely ionised at the operating pH values which are usually between 2 and 8. An equilibrium is set up between the neutral species of a mixture in the stationary phase and the respective ionised (anion or cation) species which dissolve in the mobile phase containing the counter ions. The extent of the equilibrium will depend on the ionisation constants of the respective components of the mixture, and the solubility of the unionised species in the stationary phase. Since the ionisation constants and the solubility in the stationary phase will vary with the water-methanol ratio of the mobile phase, the separation may be improved by altering this ratio gradually (gradient elution) or stepwise. If the compounds are eluted too rapidly, the water content of the mobile phase should be increased, e.g. by steps of 10%. Conversely, if components do not move, or move slowly, the methanol content of the mobile phase should be increased by steps of 10%.

The application of pressure to the liquid phase in liquid chromatography generally increases the separation (see HPLC). In PIC also, improved efficiency of the column is observed if pressure is applied to the mobile phase (Wittmer et al. *Anal Chem* **47** 1422 1975). [See the Fluka (Riedel-deHaën) catalogue and Supelco catalogue, <<http://sigma-aldrich.dirxion.com/WebProject.asp?BookCode=chr09flx#>> for IPC reagents for the separation of cations and anions, ]

### Ion-exchange Chromatography

Ion-exchange chromatography involves an electrostatic process which depends on the relative affinities of various types of ions for an immobilised assembly of ions of opposite charge. The mobile phase is an aqueous buffer with a fixed pH or an aqueous mixture of buffers in which the pH is continuously increased or decreased as the separation may require. This form of liquid chromatography can also be performed at high inlet pressures of liquid with increased column performances.

**Ion-exchange Resins.** An ion-exchange resin is made up of particles of an insoluble elastic hydrocarbon network to which is attached a large number of ionisable groups. Materials commonly used comprise synthetic ion-exchange resins made, for example, by crosslinking polystyrene to which has been attached non-diffusible ionised or ionisable groups. Resins with relatively high crosslinkage (8–12%) are suitable for the chromatography of small ions, whereas those with low cross linkage (2–4%) are suitable for larger molecules. Applications to hydrophobic systems are possible using aqueous gels with phenyl groups bound to the rigid matrix (Phenyl-Superose/Sepharose, Pharmacia-Amersham Biosciences) or neopentyl chains (Alkyl-Superose, Pharmacia-Amersham Biosciences). (Superose is a cross-linked agarose-based medium with an almost uniform bead size.) These groups are further distinguishable as strong [ $-\text{SO}_2\text{OH}$ ,  $-\text{NR}_3^+$ ] or weak [ $-\text{OH}$ ,  $-\text{CO}_2\text{H}$ ,  $-\text{PO}(\text{OH})_2$ ,  $-\text{NH}_2$ ]. Their charges are counterbalanced by diffusible ions, and the operation of a column depends on its ability and selectivity to replace these ions. The exchange that takes place is primarily an electrostatic process but adsorptive forces and hydrogen bonding can also be important. A typical sequence for the relative affinities of some common anions (and hence the inverse order in which they pass through such a column) is the following, obtained using a quaternary ammonium (strong base) anion-exchange column:

Fluoride < acetate < bicarbonate < hydroxide < formate < chloride < bromate < nitrite < cyanide < bromide < chromate < nitrate < iodide < thiocyanate < oxalate < sulfate < citrate.

For an amine (weak base) anion-exchange column in its chloride form, the following order has been observed:

Fluoride < chloride < bromide = iodide = acetate < molybdate < phosphate < arsenate < nitrate < tartrate < citrate < chromate < sulfate < hydroxide.

With strong cation-exchangers (e.g. with  $\text{SO}_3\text{H}$  groups), the usual sequence is that polyvalent ions bind more firmly than mono- or di- valent ones, a typical series being as follows:

$\text{Th}^{4+} > \text{Fe}^{3+} > \text{Al}^{3+} > \text{Ba}^{2+} > \text{Pb}^{2+} > \text{Sr}^{2+} > \text{Ca}^{2+} > \text{Co}^{2+} > \text{Ni}^{2+} = \text{Cu}^{2+} > \text{Zn}^{2+} = \text{Mg}^{2+} > \text{UO}_2^{+} = \text{Mn}^{2+} > \text{Ag}^{+} > \text{Tl}^{+} > \text{Cs}^{+} > \text{Rb}^{+} > \text{NH}_4^{+} = \text{K}^{+} > \text{Na}^{+} > \text{H}^{+} > \text{Li}^{+}$ .

Thus, if an aqueous solution of a sodium salt contaminated with heavy metals is passed through the sodium form of such a column, the heavy metal ions will be removed from the solution and will be replaced by sodium ions from the column. This effect is greatest in dilute solution. Passage of sufficiently strong solutions of alkali metal salts or mineral acids readily displaces all other cations from ion-exchange columns. (The regeneration of columns depends on this property.) However, when the cations lie well to the left in the above series it is often advantageous to use a complex-forming species to facilitate removal. For example, iron can be displaced from ion-exchange columns by passage of sodium citrate or sodium ethylenediaminetetraacetate.

Some of the more common commercially available resins are listed in Table 12.

Ion-exchange resins swell in water to an extent which depends on the amount of crosslinking in the polymer, so that columns should be prepared from the wet material by adding it as a suspension in water to a tube already partially filled with water. (This also avoids trapping air bubbles.) The exchange capacity of a resin is commonly expressed as mg equiv/ml of wet resin. This quantity is pH-dependent for weak-acid or weak-base resins but is constant at about 0.6–2 for most strong-acid or strong-base types.

Apart from their obvious applications to inorganic species, sulfonic acid resins have been used in purifying amino acids, aminosugars, organic acids, peptides, purines, pyrimidines, nucleosides, nucleotides and polynucleotides. Thus, organic bases can be applied to the  $H^+$  form of such resins by adsorbing them from neutral solution and, after washing with water, they are eluted sequentially with suitable buffer solutions or dilute acids. Alternatively, by passing alkali solution through the column, the bases will be displaced in an order that is governed by their pK values. Similarly, strong-base anion exchangers have been used for aldehydes and ketones (as bisulfite addition compounds), carbohydrates (as their borate complexes), nucleosides, nucleotides, organic acids, phosphate esters and uronic acids. Weakly acidic and weakly basic exchange resins have also found extensive applications, mainly in resolving weakly basic and acidic species. For demineralisation of solutions without large changes in pH, mixed-bed resins can be prepared by mixing a cation-exchange resin in its  $H^+$  form with an anion-exchange resin in its  $OH^-$  form. Commercial examples include Amberlite MB-1 (IR-120 + IRA-400) and Bio-Deminrolit (Zeo-Karb 225 and Zerolit FF). The latter is also available in a self-indicating form.

***Ion-exchange Celluloses and Sephadex.*** A different type of ion-exchange column that finds extensive application in biochemistry for the purification of proteins, nucleic acids and acidic polysaccharides derives from cellulose by incorporating acidic and basic groups to give ion-exchangers of controlled acid and basic strengths. Commercially available cellulose-type resins are listed in Tables 13 and 14. AG 501 x 8 (Bio-Rad) is a mixed-bed resin containing equivalents of AG 50W-x8  $H^+$  form and AG 1-x8  $HO^-$  form, and Bio-Rex MSZ 501 resin. A dye marker indicates when the resin is exhausted. Removal of unwanted cations, particularly of the transition metals, from amino acids and buffer can be achieved by passage of the solution through a column of Chelex 20 or Chelex 100. The metal-chelating abilities of the resin reside in the bonded iminodiacetate groups. Chelex can be regenerated by washing in two bed volumes of 1M HCl, two bed volumes of 1M NaOH and five bed volumes of water.

Ion-exchange celluloses are available in different particle sizes. It is important that the amounts of 'fines' are kept to a minimum otherwise the flow of liquid through the column can be extremely slow to the point of no liquid flow. Celluloses with a large range of particle sizes should be freed from 'fines' before use. This is done by suspending the powder in the required buffer and allowing it to settle for one hour and then decanting the 'fines'. This separation appears to be wasteful, but it is necessary for reasonable flow rates without applying high pressures at the top of the column. Good flow rates can be obtained if the cellulose column is packed dry whereby the 'fines' are evenly distributed throughout the column. Wet packing causes the 'fines' to rise to the top of the column, which thus becomes clogged.

Several ion-exchange celluloses require recycling before use, a process which must be applied for recovered celluloses. Recycling is done by stirring the cellulose with 0.1M aqueous sodium hydroxide, washing with water until neutral, then suspending in 0.1M hydrochloric acid and finally washing with water until neutral. When regenerating a column it is advisable to wash with a salt solution (containing the required counter ions) of increasing ionic strength up to 2M. The cellulose is then washed with water and recycled if necessary. Recycling can be carried out more than once if there are doubts about the purity of the cellulose and when the cellulose had been used previously for a different purification procedure than the one to be used. The basic matrix of these ion-exchangers is cellulose and it is important not to subject them to strong acid (> 1M) and strongly basic (> 1M) solutions.

When storing ion-exchange celluloses, or during prolonged usage, it is important to avoid growth of microorganisms or moulds which slowly destroy the cellulose. Good inhibitors of microorganisms are phenyl mercuric salts (0.001%, effective in weakly alkaline solutions), chlorohexidine (Hibitane at 0.002% for anion exchangers), 0.02% aqueous sodium azide or 0.005% of ethyl mercuric thiosalicylate (Merthiolate); these are most effective in weakly acidic solutions for cation exchangers. Trichlorobutanol (Chloretone, at 0.05% is only effective in weakly acidic solutions) can be used for both anion and cation exchangers. Most organic solvents (e.g. methanol) are effective antimicrobial agents but only at high concentrations. These inhibitors must be

removed by washing the columns thoroughly before use because they may have adverse effects on the material to be purified (e.g. inactivation of enzymes or other active preparations).

*Sephadex*. Other carbohydrate matrices such as *Sephadex* are a bead form of cross-linked gels (based on dextran) which have more uniform particle sizes. Their advantages over the celluloses include faster and more reproducible flow rates and they can be used directly without removal of 'fines'. *Sephadex*, which can also be obtained in a variety of ion-exchange forms (see Table 14) consists of beads of a cross-linked dextran gel which swells in water and aqueous salt solutions. The smaller the bead size, the higher the resolution that is possible but the slower the flow rate. Typical applications of *Sephadex* gels are the fractionation of mixtures of polypeptides, proteins, nucleic acids, polysaccharides and for desalting solutions. *Sephadex* ion-exchangers, unlike celluloses, are available in narrow ranges of particle sizes. These are of two medium types, the G-25 and G-50, and their dry bead diameter sizes are *ca* 50 to 150 microns. They are available as cation and anion exchange *Sephadex*. One of the disadvantages of using *Sephadex* ion-exchangers is that the bed volume can change considerably with alteration of pH. *Ultragels* also suffer from this disadvantage to a varying extent, but ion-exchangers of the bead type have been developed e.g. *Fractogels*, *Toyopearl*, which do not suffer from this disadvantage.

*Sepharose* (e.g. *Sepharose CL and Bio-Gel A*) is a bead form of agarose gel which is useful for the fractionation of high molecular weight substances, for molecular weight determinations of large molecules (molecular weight > 5000), and for the immobilisation of enzymes, antibodies, hormones and receptors usually for affinity chromatography applications.

In preparing any of the above for use in columns, the dry powder is evacuated, then mixed under reduced pressure with water or the appropriate buffer solution. Alternatively it is stirred gently with the solution until all air bubbles are removed. Because some of the wet powders change volumes reversibly with alteration of pH or ionic strength (see above), it is imperative to make allowances when packing columns (see above) in order to avoid overflowing of packing when the pH or salt concentrations are altered.

*Cellex CM* ion-exchange cellulose can be purified by treatment of 30–40g (dry weight) with 500ml of 1mM cysteine hydrochloride. It is then filtered through a Büchner funnel and the filter cake is suspended in 500ml of 0.05M NaCl/0.5M NaOH. This is filtered and the filter cake is resuspended in 500ml of distilled water and filtered again. The process is repeated until the washings are free from chloride ions. The filter cake is again suspended in 500ml of 0.01M buffer at the desired pH for chromatography, filtered, and the last step repeated several times.

*Cellex D* and other anionic celluloses are washed with 0.25M NaCl/0.25M NaOH solution, then twice with deionised water. This is followed with 0.25M NaCl and then washed with water until chloride-free. The *Cellex* is then equilibrated with the desired buffer as above.

*Crystalline Hydroxylapatite* is a structurally organised, highly polar material which, in aqueous solution (in buffers) strongly adsorbs macromolecules such as proteins and nucleic acids, permitting their separation by virtue of the interaction with charged phosphate groups and calcium ions, as well by physical adsorption. The procedure therefore is not entirely ion-exchange in nature. Chromatographic separations of singly and doubly stranded DNA are readily achievable, whereas there is negligible adsorption of low-molecular-weight species.

### Gel Filtration

The gel-like, bead nature of wet *Sephadex* enables small molecules such as inorganic salts to diffuse freely into it while, at the same time, protein molecules are unable to do so. Hence, passage through a *Sephadex* column can be used for complete removal of salts from protein solutions. Polysaccharides can be freed from monosaccharides and other small molecules because of their differential retardation. Similarly, amino acids can be separated from proteins and large peptides.

Gel filtration using *Sephadex* G-types (50 to 200) is essentially useful for fractionation of large molecules with molecular weights above 1000. For *Superose*, the range is given as 5000 to  $5 \times 10^6$ . Fractionation of lower molecular weight solutes (e.g. ethylene glycols, benzyl alcohols) can now be achieved with *Sephadex* G-10 (up to Mol.Wt 700) and G-25 (up to Mol.Wt 1500). These dextrans are used only in aqueous solutions. In contrast, *Sephadex* LH-20 and LH-60 (prepared by hydroxypropylation of *Sephadex*) are used for the separation of small molecules (Mol.Wt less than 500) using most of the common organic solvents as well as water.

*Sephasorb HP* (ultrafine, prepared by hydroxypropylation of cross-linked dextran) can also be used for the separation of small molecules in organic solvents and water, and in addition it can withstand pressures up to

1400 psi making it useful in HPLC. These gels are best operated at pH values between 2 and 12, because solutions with high and low pH values slowly decompose them (see further in Chapter 7).

Supelco (see catalogue) supply a variety of SUPELCOGEL columns (for small molecule separations), TSK-GEL columns (for large molecules separation) and guard columns for gel permeation chromatography. They have columns of the latter type (e.g. TSK-GEL column G4000SW) which can separate globular proteins of 20–10,000  $\times 10^3$  Daltons in molecular weight. They also supply “Ascentis HPLC Applications CDs” containing a comprehensive library of their columns and possible applications.

### High Performance Liquid Chromatography (HPLC)

When pressure is applied at the inlet of a liquid chromatographic column the performance of the column can be increased by several orders of magnitude. This is partly because of the increased speed at which the liquid flows through the column and partly because fine column packings which have larger surface areas can be used. Because of the improved efficiency of the columns, this technique has been referred to as high performance, high pressure, or high speed liquid chromatography and has found great importance in chemistry and biochemistry.

The equipment consists of a hydraulic system to provide the pressure at the inlet of the column, a column, a detector, data storage and output, usually in the form of a computer. The pressures used in HPLC vary from a few psi to 4000–5000 psi. The most convenient pressures are, however, between 500 and 1800psi. The plumbing is made of stainless steel or non-corrosive metal tubing to withstand high pressures. Plastic tubing and connectors are used for low pressures, e.g. up to ~500psi. Increase of temperature has a very small effect on the performance of a column in liquid chromatography. Small variations in temperatures, however, do upset the equilibrium of the column, hence it is advisable to place the column in an oven at ambient temperature in order to achieve reproducibility. The packing (stationary phase) is specially prepared for withstanding high pressures. It may be an adsorbent (for adsorption or solid-liquid HPLC), a material impregnated with a high boiling liquid (e.g. octadecyl sulfate, in *reverse-phase* or *liquid-liquid* or *paired-ion* HPLC), an ion-exchange material (in *ion-exchange* HPLC), or a highly porous non-ionic gel (for high performance *gel filtration* or *permeation*). The mobile phase is water, aqueous buffers, salt solutions, organic solvents or mixtures of these.

### Detectors

The more commonly used detectors for column chromatography in general have UV, visible, diode array or fluorescence monitoring for light absorbing substances in the effluent, and refractive index monitoring and evaporative light scattering for transparent compounds in the effluent. UV detection is not useful when molecules do not have UV absorbing chromophores, and solvents for elution should be carefully selected when UV monitoring is used so as to ensure the lack of background interference in detection. The sensitivity of the refractive index monitoring is usually lower than the light absorbing monitoring by a factor of ten or more. It is also difficult to use a refractive index monitoring system with gradient elution of solvents. When substances have readily oxidised and reduced forms, e.g. phenols, nitro compounds, heterocyclic compounds etc. then electrochemical detectors are useful. These detectors oxidise and/or reduce these substances and make use of this process to provide a peak on the recorder.

The cells of the monitoring devices are very small (*ca* 5  $\mu$ l) and the detection is very good. The volumes of the analytical columns are quite small (*ca* 2ml for a 1 metre column); hence the result of an analysis is achieved very quickly. Larger columns have been used for preparative work and can be used with the same equipment. Most machines have solvent mixing chambers for solvent gradient or ion gradient elution. The solvent gradient (for two solvents) or pH or ion gradient can be adjusted in a linear, increasing or decreasing exponential manner. Splitters can be used, whereby very small volumes of the effluent are directed through the detectors so that the whole effluent does not pass through the detector.

### Columns for HPLC

In general two different types of HPLC columns are available. Prepacked columns are those with metal casings with threads at both ends onto which capillary connections are attached. The cartridge HPLC columns are cheaper and are used with cartridge holders. As the cartridge is fitted with a groove for the holding device, no threads are necessary and the connection pieces can be reused.

A large range of HPLC columns (including guard columns, i.e. small pre-columns) are available from Supelco < <http://sigma-aldrich.dirxion.com/WebProject.asp?BookCode=chr09flx#> >, Waters <[www.waters.com](http://www.waters.com)>, Agilent Technologies <[www.chem.agilent.com](http://www.chem.agilent.com)>, Phenomenex <[www.phenomenex.com](http://www.phenomenex.com)>, YMC <[www.ymc.co.jp/en/](http://www.ymc.co.jp/en/)>, Merck <[www.merck.de](http://www.merck.de)>, SGE <[www.sge.com](http://www.sge.com)>, GE Healthcare

<[www.gehealthcare.com](http://www.gehealthcare.com)>, and other leading companies. It is not possible to list the range of columns here that are commercially available because the numbers are too large and include prepared columns for the type of chromatography described below in the *Other Types of Liquid Chromatography* such as Monolithic Chromatography and UPLC (see below). Also, in this range of columns are columns with chiral bonded phases capable of separating enantiomeric mixtures. The number of these, on the other hand, is relatively smaller and some *chiral* columns are listed in Table 15.

### Other Types of Liquid Chromatography

New stationary phases for specific purposes in chromatographic separation are being continually developed. *Charge transfer adsorption chromatography* makes use of a stationary phase which contains immobilised aromatic compounds and permits the separation of aromatic compounds by virtue of the ability to form charge transfer complexes (sometimes coloured) with the stationary phase. The separation is caused by the differences in stability of these complexes (Porath and Dahlgren-Caldwell *J Chromatogr* **133** 180 1977).

In *metal chelate adsorption chromatography* a metal is immobilised by partial chelation on a column which contains bi- or tri- dentate ligands. Its application is in the separation of substances which can complex with the bound metals and depends on the stability constants of the various ligands (Porath et al. *Nature* **258** 598 1975; Loennerdal et al. *FEBS Lett* **75** 89 1977).

An application of chromatography which has found extensive use in biochemistry and has brought a new dimension in the purification of enzymes is *affinity chromatography*. A specific enzyme inhibitor is attached by covalent bonding to a stationary phase (e.g. AH-Sepharose 4B for acidic inhibitors and CH-Sepharose 4B for basic inhibitors, Phenyl-Sepharose for hydrophobic proteins), and will strongly bind only the specific enzyme which is inhibited or preferentially bound, allowing all other proteins to flow through the column. The enzyme is then eluted with a solution of high ionic strength (e.g. 1M sodium chloride) or a solution containing a substrate or reversible inhibitor of the specific enzyme. (The ionic medium can be removed by gel filtration using a mixed-bed gel.) Similarly, an immobilised lectin may interact with the carbohydrate moiety of a glycoprotein. The most frequently used matrixes are cross-linked (4-6%) agarose and polyacrylamide gel. Many adsorbents are commercially available for nucleotides, coenzymes and vitamins, amino acids, peptides, lectins and related macromolecules and immunoglobulins. Considerable purification can be achieved by one passage through the column and the column can be reused several times.

The affinity method may be *biospecific*, for example as an antibody-antigen interaction, or chemical as in the chelation of boronate by *cis*-diols, or of unknown origin as in the binding of certain dyes to albumin and other proteins.

*Hydrophobic adsorption chromatography* takes advantage of the hydrophobic properties of substances to be separated and has also found use in biochemistry (Hoftsee *Biochem Biophys Res Commun* **50** 751 1973; Jennissen & Heilmayer Jr *Biochemistry* **14** 754 1975). Specific covalent binding with the stationary phase, a procedure that was called *covalent chromatography*, has been used for the separation of compounds and for immobilising enzymes on a support: the column was then used to carry out specific bioorganic reactions (Mosbach *Method Enzymol* **44** 1976; A. Rosevear et al. *Immobilised Enzymes and Cells: A Laboratory Manual*, Adam Hilger, Bristol, 1987, ISBN 085274515X). See Bibliography for further literature.

More recently *Monolithic Chromatography* has been introduced which is a new type of high-performance liquid chromatography in which the columns are a 'one-piece porous solid', or *monolith*, instead of particles. These columns take a variety of forms for use in adsorption, ion exchange (weak and strong, cation and anion), reverse phase, and are for use in the separation of small and large molecules. The mobile phase in these columns flows through the whole of the stationary phase. [P. Wang ed., *Monolithic Chromatography and its Modern Applications* ILM Publications, pp 648 2010, ISBN 9781906799038, 1906799032; and for columns see BIA Separations <[www.biaseparations.com](http://www.biaseparations.com)>.

*Ultra performance Liquid chromatography (UPLC)* affords a considerable improvement by bringing high performance liquid chromatography to a new level. Great improvements in analysis and purification of amino acids, peptides, proteins, oligonucleotides and glycans can be accomplished. This has been achieved by packing columns with smaller sized particles (1.7–1.8 $\mu$ m) and applying pressures of ~15,000psi (~1030 bar) to the mobile phase. NanoACQUITY UPLC trapping and nanoflow columns have been specifically designed for use on Waters nanoACQUITY systems that can be integrated with MS components [see [www.waters.com](http://www.waters.com)].

### Automated column chromatography

Most of the above methods of column chromatography have been, or can be, automated. Devices are available for the automated injection of samples to columns which are useful for analytical evaluation of samples, for

repeated analyses, or for repeated separations to obtain larger amounts of material. The specific fractions of the effluent can be collected. Equipment for these purposes can be obtained from several of the suppliers listed at the end of the HPLC section above with the corresponding websites. GC systems coupled with mass spectrometers (GC-MS) and HPLC systems coupled to mass spectrometers (LC-MS) are extremely important methods for the separation and identification of substances. These are invariably linked to a computer with internal libraries which can identify the peaks, and the libraries can be continually updated (see above). With more elaborate equipment LC-MS-MS where the peaks from the first spectrometer are further analysed by a second mass spectrometer provide a wealth of information. If not for the costs involved in GC-MS, GC-MS-MS, LC-MS and LC-MS-MS equipment, these systems would be more commonly found in analytical and research laboratories. [For further reading see Bibliography.]

## ELECTROPHORESIS

Ionisable substances such as organic and inorganic acids, bases and salts migrate to their respective electrodes (anode or cathode) if a voltage is applied. When they are placed onto a matrix, e.g. paper or gel, then their rate of migration to the electrodes will vary with the charge, nature and structure of the substance. This phenomenon is known as electrophoresis and is very useful for separating and purifying substances. Capillary techniques have been adapted to electrophoresis and “capillary electrophoresis”, and “capillary zone electrophoresis” are finding wide use for identification, separation and isolation of ionisable substances (see text in the Bibliography under “electrophoresis” and the “Introduction” in Chapter 7). The method is used extensively for biological substances, e.g. proteins, polypeptides, DNA, RNA, (see Introduction in Chapter 7) but has been used to a limited extent for identifying and purifying small molecules. Elaborate equipment is available commercially which contains essentially an electrolytic cell and a power supply which provides variable voltage for the process. The use of paper (Whatman of various thicknesses) as the matrix on a flat bed or in a vertical descending mode has been completely superseded with polyacrylamide or agarose flat bed gels. These are routinely used mainly for the separation of proteins and nucleic acids. Also capillary electrophoresis (CE) is now widely used for the analysis and detection of biological substances. It is used for the separation and purification of carbohydrates, nucleic acids, proteins and peptides and for chiral analysis and separations [see Bibliography].

## DRYING

### Removal of Solvents

Where substances are sufficiently stable, removal of solvent from recrystallised materials presents no problems. The crystals, after filtering at the pump (and perhaps air-drying by suction), are heated in an oven above the boiling point of the solvent (but below the melting point of the crystals), followed by cooling in a desiccator. Where this treatment is inadvisable, it is still often possible to heat to a lower temperature under reduced pressure, for example in an Abderhalden pistol. This device consists of a small chamber which is heated externally by the vapour of a boiling solvent. Inside this chamber, which can be evacuated (pump) is placed a small boat containing the sample to be dried and also a receptacle with a suitable drying agent. Convenient liquids for use as boiling liquids in an Abderhalden pistol, and their boiling temperatures, are given in Table 16. Alternatively an electrically heated drying pistol can also be used. In cases where heating above room temperature cannot be used, drying must be carried out in a vacuum desiccator containing suitable absorbents. For example, hydrocarbons, such as cyclohexane and petroleum ether, can be removed by using shredded paraffin wax, and acetic acid and other acids can be absorbed by pellets of sodium or potassium hydroxide. However, in general, solvent removal is less of a problem than ensuring that the water content of solids and liquids is reduced below an acceptable level.

### Removal of Water

Methods for removing water from solids depend on the thermal stability of the solids or the time available. The safest way is to dry in a vacuum desiccator over concentrated sulfuric acid, phosphorus pentoxide, silica gel, calcium chloride, or some other desiccant. Where substances are stable in air and melt above 100°, drying in an air oven may be adequate. In other cases, use of an Abderhalden pistol may be satisfactory. Often, in drying inorganic salts, the final material that is required is a hydrate. In such cases, the purified substance is left in a desiccator to equilibrate above an aqueous solution having a suitable water-vapour pressure. A convenient range of solutions used in this way is given in Table 17. The choice of desiccants for drying liquids is more restricted because of the need to avoid all substances likely to react with the liquids themselves. In some cases, direct distillation of an organic liquid is a suitable method

for drying both solids and liquids, especially if low-boiling azeotropes are formed. Examples include acetone, aniline, benzene, chloroform, carbon tetrachloride, heptane, hexane, methanol, nitrobenzene, petroleum ether, toluene and xylene. Addition of benzene can be used for drying ethanol by distillation. In carrying out distillations intended to yield anhydrous products, the apparatus should be fitted with guard-tubes containing calcium chloride or silica gel to prevent entry of moist air into the system. (Many anhydrous organic liquids are appreciably hygroscopic.)

Traces of water can be removed from solvents such as benzene, 1,2-dimethoxyethane, diethyl ether, pentane, toluene and tetrahydrofuran by refluxing under nitrogen a solution containing sodium wire and benzophenone, and fractionally distilling. Drying with, and distilling from  $\text{CaH}_2$  is applicable to a number of solvents including aniline, benzene, *tert*-butylamine, *tert*-butanol, 2,4,6-collidine, diisopropylamine, dimethylformamide, hexamethyl-phosphoramide, dichloromethane, ethyl acetate, pyridine, tetramethylethylene diamine, toluene, triethylamine.

Removal of water from gases may be by physical or chemical means, and is commonly by adsorption on to a drying agent in a low-temperature trap. The effectiveness of drying agents depends on the vapour pressure of the hydrated compound - the lower the vapour pressure the less the remaining moisture in the gas.

The most usually applicable of the specific methods for detecting and determining water in organic liquids is due to Karl Fischer. (See J. Mitchell & D.M. Smith, *Aquametry*, 2nd Ed, J Wiley & Sons, New York, 1977–1984, ISBN 0471022640; Fieser & Fieser, *Reagents for Organic Synthesis*, J.Wiley & Sons, NY, Vol 1, 528 1967, ISBN 0271616X), also see Karl Fischer titrant or Hydranal<sup>®</sup> –Titrant Type 5E [64-17-5] and other types in the Fluka and <<http://www.sigmaaldrich.com/analytical-chromatography/titration/hydranal.html>>. Other techniques include electrical conductivity measurements and observation of the temperature at which the first cloudiness appears as the liquid is cooled (applicable to liquids in which water is only slightly soluble). Addition of anhydrous cobalt (II) iodide (blue) provides a convenient method (colour change to pink on hydration) for detecting water in alcohols, ketones, nitriles and some esters. Infrared absorption measurements of the broad band for water near  $3500\text{ cm}^{-1}$  can also sometimes be used for detecting water in non-hydroxylic substances.

Cartridges for the removal not only water from solvents or solutions but other specific impurities, e.g. acids, amines, aldehydes, are now commercially available [see supplies listed at the end of the HPLC section together with their respective websites].

For further useful information on mineral adsorbents and drying agents, go to the SigmaAldrich website <[sigmaaldrich.com](http://sigmaaldrich.com)>, under technical library (Aldrich) for technical bulletin AL-143.

### Intensity and Capacity of Common Desiccants

Drying agents are conveniently grouped into three classes, depending on whether they combine with water reversibly, they react chemically (irreversibly) with water, or they are molecular sieves. The first group varies in their drying intensity with the temperature at which they are used, depending on the vapour pressure of the hydrate that is formed. This is why, for example, drying agents such as anhydrous sodium sulfate, magnesium sulfate or calcium chloride should be filtered off from the liquids before the latter are heated. The intensities of drying agents belonging to this group fall in the sequence:

$\text{P}_2\text{O}_5 \gg \text{BaO} > \text{Mg}(\text{ClO}_4)_2, \text{CaO}, \text{MgO}, \text{KOH} \text{ (fused)}, \text{conc H}_2\text{SO}_4, \text{CaSO}_4, \text{Al}_2\text{O}_3 > \text{KOH} \text{ (pellets)},$   
 silica gel,  $\text{Mg}(\text{ClO}_4)_2 \cdot 3\text{H}_2\text{O} > \text{NaOH} \text{ (fused)}, 95\% \text{ H}_2\text{SO}_4, \text{CaBr}_2, \text{CaCl}_2 \text{ (fused)} > \text{NaOH} \text{ (pellets)},$   
 $\text{Ba}(\text{ClO}_4)_2, \text{ZnCl}_2, \text{ZnBr}_2 > \text{CaCl}_2 \text{ (technical)} > \text{CuSO}_4 > \text{Na}_2\text{SO}_4, \text{K}_2\text{CO}_3.$

Where large amounts of water are to be removed, a preliminary drying of liquids is often possible by shaking with concentrated solutions of sodium sulfate or potassium carbonate, or by adding sodium chloride to salt out the organic phase (for example, in the drying of lower alcohols), as long as the drying agent does not react (e.g.  $\text{CaCl}_2$  with alcohols and amines, see below).

Drying agents that combine irreversibly with water include the alkali metals, the metal hydrides (discussed in Chapter 2), and calcium carbide.

### Suitability of Individual Desiccants

**Alumina.** (Preheated to  $175^\circ$  for about 7 hours). Mainly as a drying agent in a desiccator or as a column through which liquid is percolated.

**Aluminium amalgam.** Mainly used for removing traces of water from alcohols *via* refluxing followed by distillation.

**Barium oxide.** Suitable for drying organic bases.

**Barium perchlorate.** Expensive. Used in desiccators (*covered with a metal guard*). Unsuitable for drying solvents or organic material where contact is necessary, because of the danger of **EXPLOSION**.

**Boric anhydride.** (Prepared by melting boric acid in an air oven at a high temperature, cooling in a desiccator, and powdering.) Mainly used for drying formic acid.

**Calcium chloride (anhydrous).** Cheap. Large capacity for absorption of water, giving the hexahydrate below 30°, but is fairly slow in action and not very efficient. Its main use is for preliminary drying of alkyl and aryl halides, most esters, saturated and aromatic hydrocarbons and ethers. Unsuitable for drying alcohols and amines (which form addition compounds), fatty acids, amides, amino acids, ketones, phenols, or some aldehydes and esters. Calcium chloride is suitable for drying the following gases: hydrogen, hydrogen chloride, carbon monoxide, carbon dioxide, sulfur dioxide, nitrogen, methane, oxygen, also paraffins, ethers, olefins and alkyl chlorides.

**Calcium hydride.** See Chapter 2.

**Calcium oxide.** (Preheated to 700–900° before use.) Suitable for alcohols and amines (but does not dry them completely). Need not be removed before distillation, but in that case the head of the distillation column should be packed with glass wool to trap any calcium oxide powder that might be carried over. Unsuitable for acidic compounds and esters. Suitable for drying gaseous amines and ammonia.

**Calcium sulfate (anhydrous).** (Prepared by heating the dihydrate or the hemihydrate in an oven at 235° for 2–3 hours; it can be regenerated.) Available commercially as *Drierite*. It forms the hemihydrate,  $2\text{CaSO}_4 \cdot \text{H}_2\text{O}$ , so that its capacity is fairly low (6.6% of its weight of water), and hence is best used on partially dried substances. It is very efficient (being comparable with phosphorus pentoxide and concentrated sulfuric acid). Suitable for most organic compounds. Solvents boiling below 100° can be dried by direct distillation from calcium sulfate.

**Copper (II) sulfate (anhydrous).** Suitable for esters and alcohols. Preferable to sodium sulfate in cases where solvents are sparingly soluble in water (for example, benzene or toluene). The colourless to fawn coloured powder turns blue as it absorbs water

**Lithium aluminium hydride.** See Chapter 2.

**Magnesium amalgam.** Mainly used for removing traces of water from alcohols by refluxing the alcohol in the presence of the Mg amalgam followed by distillation.

**Magnesium perchlorate (anhydrous).** (Available commercially as *Dehydrite*. Expensive.) Used in desiccators. Unsuitable for drying solvents or any organic material where contact is necessary, because of the danger of **EXPLOSION**.

**Magnesium sulfate (anhydrous).** (Prepared from the heptahydrate by drying at 300° under reduced pressure.) More rapid and effective than sodium sulfate but is slightly acidic. It has a large capacity, forming  $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$  below 48°. Suitable for the preliminary drying of most organic compounds.

**Molecular sieves.** See below.

**Phosphorus pentoxide.** Very rapid and efficient, but difficult to handle and should only be used after the organic material has been partially dried, for example with magnesium sulfate. Suitable for anhydrides, alkyl and aryl halides, ethers, esters, hydrocarbons and nitriles, and for use in desiccators. Not suitable with acids, alcohols, amines or ketones, or with organic molecules from which a molecule of water can be eliminated. Suitable for drying the following gases: hydrogen, oxygen, carbon dioxide, carbon monoxide, sulfur dioxide, nitrogen, methane, ethene and paraffins. It is available on a solid support with an indicator under the name *Sicapent* (from Merck). The colour changes in Sicapent depend on the percentage of water present (e.g. in the absence of water, Sicapent is colourless but becomes green with 20% water and blue with 33% w/w water). When the quantity of water in the desiccator is high, a crust of phosphoric acid forms a layer over the phosphorus pentoxide powder and decreases its efficiency. The crust can be removed with a spatula to expose the dry powder and restore the desiccant property.

**Potassium (metal).** Properties and applications are similar to those for sodium but as the reactivity is greater than that of sodium, the hazards are greater than those of sodium. **Handle with extreme care.**

**Potassium carbonate (anhydrous).** Has a moderate efficiency and capacity, forming the dihydrate. Suitable for an initial drying of alcohols, bases, esters, ketones and nitriles by shaking with them, then filtering off. Also suitable for salting out water-soluble alcohols, amines and ketones. Unsuitable for acids, phenols, thiols and other acidic substances.

**Potassium carbonate.** Solid potassium hydroxide is very rapid and efficient. Its use is limited almost entirely to the initial drying of organic bases. Alternatively, sometimes the base is shaken first with a concentrated solution of potassium hydroxide to remove most of the water present. Unsuitable for acids, aldehydes, ketones, phenols, thiols, amides and esters. Also used for drying gaseous amines and ammonia.

**Silica gel.** Granulated silica gel is a commercially available drying agent for use with gases, in desiccators, and (because of its chemical inertness) in physical instruments (pH meters, spectrometers, balances). Its drying action depends on physical adsorption, so that silica gel must be used at room temperature or below. By incorporating cobalt chloride into the material it can be made self indicating (blue when dry, pink when wet), re-drying in an oven at 110° being necessary when the colour changes from blue to pink.

**Sodium (metal).** Used as a fine wire or as chips, for more completely drying ethers, saturated hydrocarbons and aromatic hydrocarbons which have been partially dried (for example with calcium chloride or magnesium sulfate). Unsuitable for acids, alcohols, alkyl halides, aldehydes, ketones, amines and esters. Reacts violently if water is present and can cause a fire with highly flammable liquids.

**Sodium hydroxide.** Properties and applications are similar to those for potassium hydroxide.

**Sodium-potassium alloy.** Used as lumps. Lower melting than sodium, so that its surface is readily renewed by shaking. Properties and applications are similar to those for sodium.

**Sodium sulfate (anhydrous).** Has a large capacity for absorption of water, forming the decahydrate below 33°, but drying is slow and inefficient, especially for solvents that are sparingly soluble in water. It is suitable for the preliminary drying of most types of organic compounds.

**Sulfuric acid (concentrated).** Widely used in desiccators. Suitable for drying bromine, saturated hydrocarbons, alkyl and aryl halides. Also suitable for drying the following gases: hydrogen, nitrogen, carbon dioxide, carbon monoxide, chlorine, methane and paraffins. Unsuitable for alcohols, bases, ketones or phenols. Also available on a solid support with an indicator under the name *Sicacide* (from Merck) for desiccators. The colour changes in *Sicacide* depends on the percentage of water present (e.g. when dry *Sicacide* is red-violet but becomes pale violet with 27% water and pale yellow to colourless with 33% w/w water).

*For convenience, many of the above drying agents are listed in Table 18 under the classes of organic compounds for which they are commonly used.*

### Molecular sieves

Molecular sieves are types of adsorbents composed of crystalline zeolites (sodium and calcium aluminosilicates). By heating them, water of hydration is removed, leaving holes of molecular dimensions in the crystal lattices. These holes are of uniform size and allow the passage into the crystals of small molecules, but not of large ones.

This *sieving* action explains their use as very efficient drying agents for gases and liquids. The pore size of these sieves can be modified (within limits) by varying the cations built into the lattices. The four types of molecular sieves currently available are:

**Type 3A sieves.** A crystalline potassium aluminosilicate with a pore size of about 3 Angstroms. This type of molecular sieves is suitable for drying liquids such as acetone, acetonitrile, methanol, ethanol and 2-propanol, and drying gases such as acetylene, carbon dioxide, ammonia, propylene and butadiene. The material is supplied as beads or pellets.

**Type 4A sieves.** A crystalline sodium aluminosilicate with a pore size of about 4 Angstroms, so that, besides water, ethane molecules (but not butane) can be adsorbed. This type of molecular sieves is suitable for drying chloroform, dichloromethane, diethyl ether, dimethylformamide, ethyl acetate, cyclohexane, benzene, toluene, xylene, pyridine and diisopropyl ether. It is also useful for low pressure air drying. The material is supplied as beads, pellets or powder.

**Type 5A sieves.** A crystalline calcium aluminosilicate with a pore size of about 5 Angstroms, these sieves adsorb larger molecules than type 4A. For example, as well as the substances listed above, propane, butane, hexane, butene, higher *n*-olefins, *n*-butyl alcohol and higher *n*-alcohols, and cyclopropane can be adsorbed, but not branched-chain C<sub>6</sub> hydrocarbons, cyclic hydrocarbons such as benzene and cyclohexane, or secondary and tertiary alcohols, carbon tetrachloride or boron trifluoride. This is the type generally used for drying gases, though organic liquids such as THF and dioxane can be dried with this type of molecular sieves.

**Type 13X sieves.** A crystalline sodium aluminosilicate with a pore size of about 10 Angstroms which enables many branched-chain and cyclic compounds to be adsorbed, in addition to all the substances removed by type 5A sieves.

They are unsuitable for use with strong acids but are stable over the pH range 5–11.

Because of their selectivity, molecular sieves offer advantages over silica gel, alumina or activated charcoal, especially in their very high affinity for water, polar molecules and unsaturated organic compounds. Their relative efficiency is greatest when the impurity to be removed is present at low concentrations. Thus, at 25° and a relative humidity of 2%, type 5A molecular sieves adsorb 18% by weight of water, whereas for silica gel and alumina the figures are 3.5 and 2.5% respectively. Even at 100° and a relative humidity of 1.3%, molecular sieves adsorb about 15% by weight of water.

The greater preference of molecular sieves for combining with water molecules explains why this material can be used for drying ethanol and why molecular sieves are probably the most universally useful and efficient drying agents. Percolation of ethanol with an initial water content of 0.5% through a 144 cm long column of type 4A molecular sieves reduced the water content to 10ppm. Similar results have been obtained with pyridine.

The main applications of molecular sieves to purification comprise:

1. Drying of gases and liquids containing traces of water.
2. Drying of gases at elevated temperatures.
3. Selective removal of impurities (including water) from gas streams.

(For example, carbon dioxide from air or ethene; nitrogen oxides from nitrogen; methanol from diethyl ether. In general, carbon dioxide, carbon monoxide, ammonia, hydrogen sulfide, mercaptans, ethane, ethene, acetylene (ethyne), propane and propylene are readily removed at 25°. In mixtures of gases, the more polar ones are preferentially adsorbed).

The following applications include the removal of straight-chain from branched-chain or cyclic molecules. For example, type 5A sieves will adsorb *n*-butyl alcohol but not its branched-chain isomers. Similarly, it separates *n*-tetradecane from benzene, or *n*-heptane from methylcyclohexane.

The following liquids have been dried with molecular sieves: acetone, acetonitrile, acrylonitrile, allyl chloride, amyl acetate, benzene, butadiene, *n*-butane, butene, butyl acetate, *n*-butylamine, *n*-butyl chloride, carbon tetrachloride, chloroethane, 1-chloro-2-ethylhexane, cyclohexane, dichloromethane, dichloroethane, 1,2-dichloropropane, 1,1-dimethoxyethane, dimethyl ether, 2-ethylhexanol, 2-ethylhexylamine, *n*-heptane, *n*-hexane, isoprene, isopropyl alcohol, diisopropyl ether, methanol, methyl ethyl ketone, oxygen, *n*-pentane, phenol, propane, *n*-propyl alcohol, propylene, pyridine, styrene, tetrachloroethylene, toluene, trichloroethylene and xylene. In addition, the following gases have been dried: acetylene, air, argon, carbon dioxide, chlorine, ethene, helium, hydrogen, hydrogen chloride, hydrogen sulfide, nitrogen, oxygen and sulfur hexafluoride.

After use, molecular sieves can be regenerated by heating at between 300–350° for several hours, preferably in a stream of dry inert gas such as nitrogen or preferably under vacuum, then cooling in a desiccator. Special precautions must be taken before regeneration of molecular sieves used in the drying of flammable solvents.

However, care must be exercised in using molecular sieves for drying organic liquids. Appreciable amounts of impurities were *formed* when samples of acetone, 1,1,1-trichloroethane and methyl-*t*-butyl ether were dried in the liquid phase by contact with molecular sieves 4A (Connett *Lab Pract* 21 545 1972). Other, less reactive types of sieves may be more suitable but, in general, it seems desirable to make a preliminary test to establish that no unwanted reaction takes place. Useful comparative data for Type 4A and 5A sieves are in Table 19. With the advent of nanotechnology, nanoparticles are finding use as porous materials for a variety of purposes [see J.A. Schwartz & C. Contescu (Eds), *Surfaces of Nanoparticles & Porous Materials*, Marcel Dekker Inc, 1999. ISBN 9780824719333].

## PROPERTIES USEFUL IN PURIFICATION

### Spectroscopic

Spectroscopic instruments of one sort or another are generally available in laboratories and useful for providing some idea of the purity of the specimen in question. Among these are IR, UV-VIS, fluorescence, NMR and mass spectrometers.

*Infrared* spectra [IR or FT(Fourier Transformed)-IR with frequency range of  $\nu$  from ~ 600 to 3400  $\text{cm}^{-1}$ ], generally of the solid grounded in a large excess of KBr, or in a mull by grinding into an oil, e.g. Nujol, or in solution, e.g.  $\text{CHCl}_3$ , provide a 'fingerprint' of the substance. The KBr spectrum, or the spectrum of a film between NaCl plates if the substance is a liquid, are more useful as they give detailed information without interfering signals from Nujol or solvent which may mask important signals. Since the IR spectra consist of several signals many of which are sharp, impurities show up clearly. However, if the impurities are less than say 10% it may be difficult to say how impure the sample is, or what impurities are present in it. On the other hand, if the sample is very pure then its spectrum will be superimposable on that of the pure authentic sample.

*Ultraviolet* Spectra (with wavelength range of  $\lambda$  from ~ 200 to 400  $\mu\text{m}$ ) are measured in dilute solution and are generally broad bands. Although the broadness of the bands make it difficult to identify impurities, the values of the molecular absorption extinction coefficients  $\epsilon$  ( $\text{M}^{-1}\text{cm}^{-1}$ ) at all wavelengths, but usually measured at the peaks or troughs, are characteristics of the substance in the particular solvent used, and would be different if the sample was impure. Glass cuvettes cannot be used as they are not transparent to UV radiation, and quartz cells should be used. However, quartz cells need only be on the two opposite faces of the four sided cuvette through which the light passes; the other two faces being made of glass. Similarly in the *visible* spectra (wavelength range of  $\lambda$  from ~ 400 to 800  $\mu\text{m}$ ) the  $\epsilon$  values are characteristics of the substance in the solvent used. In this case the cheaper glass cuvettes may be used as they are transparent to visible light.

*Fluorescence* Spectra are measured in the wavelength range similar to the *visible* range, but from light that is scattered at right angles to the incident excitation (UV) wavelength. Thus at a set excitation wavelength  $\lambda_{\text{ex}}$ , the fluorescence spectrum is scanned and the peak maximum  $\lambda_{\text{em}}$  and its  $\epsilon_{\text{em}}$  are recorded. In this case the cuvettes generally have quartz faces on all four sides, as the UV light has to go through adjacent sides of the cell. This spectroscopy is useful as sometimes impurities in a sample may fluoresce at a particular wavelength. Generally, dyes have fluorescent properties and are identified in this way. Substances with strongly fluorescing properties have found considerable use in biology. Here they have been tagged to biological molecules and their movement into particular tissues and cells has been traced through their fluorescence. Table 20 lists a number

of *Fluorochromes* which have found many applications in analytical chemistry (by tagging to non-fluorescent compounds) and in biology. By selecting a mixture of two fluorochromes it is possible to obtain a desired emission wavelength. In this case the emitted fluorescence energy from the excitation of the first fluorochrome is transferred to the second fluorochrome to provide the desired fluorescence.

For other than macromolecules it is important that *at least* the  $^1\text{H}$  NMR spectrum and/or the mass spectrum of the substance should be measured routinely. These measurements require no more than one to three milligrams of material and provide a considerable amount of information about the substance. The  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra are measured to assess the purity of hydrogen and carbon containing samples. The use of very high magnetic field NMR spectrometers is especially useful for detecting impurities in such samples. The signals and their relative heights can provide valuable information not only about the extent of the impurities, but also some indications about the nature of the impurities. A variety of NMR solvents are available for dissolving the samples, and the hydrogen atoms of the solvents are replaced by deuterium which does not interfere with the  $^1\text{H}$  or  $^{13}\text{C}$  spectra. However, deuteration is generally just under 100% and signals from residual H in the solvent may appear in the spectrum and need to be identified. Similarly  $^{13}\text{C}$  signals from solvents also should be identified. Common solvents and reagents that contain trace impurities will also show minor signals in the NMR spectra. The  $^1\text{H}$  NMR signals of trace impurities in some common organic solvents (including water) and some reagents are presented in Tables 21 and 22. Similarly presented in Table 23 are the  $^{13}\text{C}$  NMR signals of some common solvents and reagents. In some instances these minor signals have been very useful as internal standards for reporting the chemical shifts of substances, thus avoiding contamination from other added standards, particularly if the samples need to be used for further studies. The NMR spectra of other nuclei such as  $^{11}\text{B}$  and  $^{31}\text{P}$  are currently also measured routinely for boron and phosphorus containing compounds. Since the compounds invariably have only a small number of these atoms in their molecules, boron or phosphorus containing impurities are readily identified in the  $^{11}\text{B}$  or  $^{31}\text{P}$  NMR spectra.

References in the bibliography at the end of this chapter to the Aldrich-Sigma catalogues of NMR, IR and mass spectral data for a large number of the compounds are listed. These collections of spectra are extremely useful for identifying compounds and impurities. If the material appears to have several impurities, these spectra should be valuable for identifying the impurities as much as possible. Preliminary chromatographic (e.g. TLC) and spot tests could be devised to monitor the material and its impurities. Purification methods can then be devised to remove these impurities, and a monitoring method will have already been established.

### Ionisation Constants — pK

When substances ionize, their neutral species produce positive and negative species. The ionisation constants are those constant values (equilibrium constants) for the equilibria between the charged species and the neutral species, or species with a larger number of charges (e.g. between mono and dications). These ionisation constants are given as **pK** values where **pK =  $-\log K$** , and **K** is the dissociation constant for the equilibrium between the species [Albert and Serjeant, *The Determination of Ionisation Constants*, A Laboratory Manual, 3rd Edition, Chapman & Hall, New York, London, 1984, ISBN 0412242907].

The advantage of using pK values (instead of K values) is that theory (and practice) states that the pK values of ionisable substances are numerically equal to the pH of the solution at which the concentrations of ionised and neutral species are equal. For example acetic acid has a pK<sup>25</sup> value of 4.76 at 25° in H<sub>2</sub>O; then at pH 4.76 the aqueous solution contains equal amounts of acetic acid [AcOH] and acetate anion [AcO<sup>-</sup>], i.e. [AcOH]/[AcO<sup>-</sup>] of 50/50. At pH 5.76 (pK + 1) the solution contains [AcOH]/[AcO<sup>-</sup>] of 10/90, at pH 6.76 (pK + 2) the solution contains [AcOH]/[AcO<sup>-</sup>] of 1/99 etc; conversely at pH 3.76 (pK - 1) the solution contains [AcOH]/[AcO<sup>-</sup>] of 90/10, and at pH 2.76 (pK - 2) the solution contains [AcOH]/[AcO<sup>-</sup>] of 99/1.

One can readily appreciate the usefulness of pK value in purification procedures, e.g. as when purifying acetic acid. If acetic acid is placed in aqueous solution and the pH adjusted to 7.76 {[AcOH]/[AcO<sup>-</sup>] with a ratio of 0.1/99.9}, and extracted with say diethyl ether, neutral impurities will be extracted into diethyl ether leaving almost all the acetic acid in the form of AcO<sup>-</sup> in the aqueous solution. If then the pH of the solution is adjusted to 1.67 where the acid is almost all in the form AcOH, almost all of it will be extracted into diethyl ether.

Aniline will be used as a second example. It has a  $pK^{25}$  of 4.60 at 25° in  $H_2O$ . If it is placed in aqueous solution at pH 1.60 it will exist almost completely (99.9%) as the anilinium cation. This solution can then be extracted with solvents e.g. diethyl ether to remove neutral impurities. The pH of the solution is then adjusted to 7.60 whereby aniline will exist as the free base (99.9%) and can be extracted into diethyl ether in order to give purer aniline.

See Table 24 for the pH values of selected buffers.

A knowledge of the  $pK$  allows the adjustment of the pH without the need of large excesses of acids or base. In the case of inorganic compounds, knowledge of the  $pK$  is useful for adjusting the ionic species for making metal complexes which could be masked or extracted into organic solvents [Perrin and Dempsey, *Buffers for pH and Metal ion Control*, Chapman & Hall, New York, London, 1974, ISBN 0412117002], or for obtaining specific anionic species in solution e.g.  $H_2PO_4^-$ ,  $HPO_4^{2-}$  or  $PO_4^{3-}$ .

The  $pK$  values that have been entered in Chapters 4, 5 and 6 have been collected directly from the literature or from compilations of literature values for organic bases [Perrin, *Dissociation Constants of Organic Bases in Aqueous Solution*, Butterworths, London, 1965, Supplement 1972, ISBN 040870408X; Albert and Serjeant, *The Determination of Ionisation Constants*, A Laboratory Manual, 3rd Edition, Chapman & Hall, London, New York, 1984, ISBN 0412242907]; organic acids [Kortum, Vogel and Andrussov, *Dissociation Constants of Organic Acids in Aqueous Solution*, Butterworth, London, 1961; Serjeant and Dempsey, *Dissociation Constants of Organic Acids in Aqueous Solution*, Pergamon Press, Oxford, New York, 1979, ISBN 0080223397; and inorganic acids and bases [Perrin, *Ionisation Constants of Inorganic Acids and Bases in Aqueous Solution*, Second Edition, Pergamon Press, Oxford, New York, 1982, ISBN 0080292143]. Where literature values were not available, values have been predicted and assigned  $pK_{Est}$ . Most predictions should be so close to true values as to make very small difference for the purposes intended in this book. The success of the predictions, i.e. how close to the true value, depends on the availability of  $pK$  values for closely related compounds because the effect of substituents or changes in structures are generally additive [Perrin, Dempsey and Serjeant, *pKa Prediction for Organic Acids and Bases*, Chapman & Hall, London, New York, 1981, ISBN 041222190X].

**All the  $pK$  values in this book are  $pK_a$  values, the acidic  $pK$ , i.e. dissociation of  $H^+$  from an acid (AH) or from a conjugate base ( $BH^+$ ).** Occasionally  $pK_b$  values are reported in the literature but these can be converted using the equation  $pK_a + pK_b = 14$ . For strong acids e.g. sulfuric acid, and strong bases, e.g. sodium hydroxide, the  $pK$  values lie beyond the 1 to 11 pH scale and have to be measured in strong acidic and basic media. In these cases appropriate scales e.g. the  $H_0$  (for acids) and  $H_-$  (for bases) have been used [see Katritzky & Waring, *J Chem Soc* 1540 1962]. These values will be less than 1 (and negative) for acids and >11 for bases. They are rough guides to the strengths of acids and bases. Errors in the stated  $pK$  and  $pK_{Est}$  values can be judged from the numerical values given. Thus  $pK$  values of 4.55, 4.5 and 4 mean that the respective errors are better than  $\pm 0.05$ ,  $\pm 0.3$  and  $\pm 0.5$ . Values taken from the literature are written as  $pK$ , and all the values that were estimated because they were not found in the literature are written as  $pK_{Est}$ .

### **$pK$ and Temperature**

The temperatures at which the literature measurements were made are given as superscripts, e.g.  $pK^{25}$ . Where no temperature is given, it is assumed that the measurements were carried out at room temperature, e.g. 15–25°. No temperature is given for estimated values ( $pK_{Est}$ ), and these have been calculated from data at room temperature. The variation of  $pK$  with temperature is given by the equation:

$$-\delta(pK)/\delta T = (pK + 0.052\Delta S^0)/T$$

where  $T$  is in degrees Kelvin and  $\Delta S^0$  is in Joules  $deg^{-1} mol^{-1}$ . The  $\delta(pK)/dT$  in the range of temperatures between 5 to 70° is generally small (e.g. between  $\sim 0.0024$  and  $\sim 0.04$ ), and for chemical purification purposes is not a seriously deterring factor. It does, however, vary with the compound under study because  $\Delta S^0$  varies from compound to compound. The following are examples of the effect of temperature on  $pK$  values: for imidazole the  $pK$  values are 7.57 (0°), 7.33 (10°), 7.10 (20°), 6.99 (25°), 6.89 (30°), 6.58 (40°) and 6.49 (50°), and for 3,5-di-nitrobenzoic acid they are 2.60 (10°), 2.73 (20°), 2.85 (30°), 2.96 (40°) and 3.07 (40°), and for *N*-acetyl- $\beta$ -

alanine they are 4.4788 (5°), 4.4652 (10°), 4.4564 (15°), 4.4488 (20°), 4.4452 (25°), 4.4444 (30°), 4.4434 (35°) and 4.4412 (40°).

#### pK and solvent

All stated pK values in this book are for data in dilute aqueous solutions unless otherwise stated, although the dielectric constants, ionic strengths of the solutions and the method of measurement, e.g. potentiometric, spectrophotometric etc., are not given. Estimated values are also for dilute aqueous solutions whether or not the material is soluble enough in water. Generally the more dilute the solution the closer is the pK to the real thermodynamic value. The pK in mixed aqueous solvents can vary considerably with the relative concentrations and with the nature of the solvents. For example the pK<sup>25</sup> values for *N*-benzylpenicillin are 2.76 and 4.84 in H<sub>2</sub>O and H<sub>2</sub>O/EtOH (20:80) respectively; the pK<sup>25</sup> values for (–)-ephedrine are 9.58 and 8.84 in H<sub>2</sub>O and H<sub>2</sub>O/MeOCH<sub>2</sub>CH<sub>2</sub>OH (20:80), respectively; and for cyclopentylamine the pK<sup>25</sup> values are 10.65 and 4.05 in H<sub>2</sub>O and H<sub>2</sub>O/EtOH (50:50) respectively. pK values in acetic acid or aqueous acetic acid are generally lower than in H<sub>2</sub>O.

The dielectric constant of the medium affects the equilibria where charges are generated in the dissociations e.g.  $AH \rightleftharpoons A^- + H^+$  and therefore affects the pK values. However, its effect on dissociations where there are no changes in total charge such as  $BH^+ \rightleftharpoons B + H^+$  is considerably less, with a slight decrease in pK with decreasing dielectric constant.

#### Solubilities of Gases in Liquids

There are two ways to define the solubilities of gases in water.

The first is the *Bunsen coefficient* ( $\beta$ ), which is the ratio of the volume of gas corrected to STP (0°C and 1atm, i.e. 760 mmHg) that dissolves in unit volume of solvent at the temperature of the experiment in equilibrium with the gas at 1atm. The second is the *Ostwald coefficient* ( $I$ ) which is the ratio of the volume of gas that dissolves in unit volume of solvent at the temperature of the experiment in equilibrium with the gas at 1atm.

The latter is a more convenient ratio to use because no correction for volume is required. Note that the volume of an ideal gas occupied by one molecular weight in grams of element or compound is ~22.4L at STP (e.g. 32g of oxygen occupy 22.4L at STP). The discussion will be limited to the solubilities of oxygen, nitrogen and air (which behave almost as ideal gases) in water, water containing salts, and in some organic solvents. Generally the solubility of these three gases in water decreases with increase of temperature and can be “boiled out” of the liquid. Their solubilities in organic liquids, on the other hand, generally increase with increase of temperature. The presence of salts in water tends to decrease the solubilities of these gases, i.e. a salting out effect, and increase in pressure increases their solubilities. These properties have to be noted in liquid chromatography at atmospheric and at high pressures. They become important when purifying small amounts of compounds by crystallisation or chromatography when large amounts of solvents are used. One must be wary of the presence of oxygen in solution, particularly in the presence of organic matter. Also the formation of reactive oxygen species e.g. “singlet” oxygen, superoxide and hydroxyl radicals, especially in the presence of trace metals such as iron, and/or of ultraviolet light can result in the formation of impurities.

The composition of air is: 78.08% of N<sub>2</sub>, 20.95% of O<sub>2</sub>, 0.03% of CO<sub>2</sub>, 0.93% of Ar and less than 0.01% of other gases. Although the partial pressure of O<sub>2</sub> in air at 1atm is ~0.20, it has a higher solubility in H<sub>2</sub>O than N<sub>2</sub>. At STP the solubility of O<sub>2</sub> by volume in H<sub>2</sub>O is 34.9% when in equilibrium with excess of air. Thus by successively dissolving air in H<sub>2</sub>O, expelling it, and redissolving the expelled air six to seven times it is possible to increase the concentration of oxygen by volume in the expelled air to 90%. The ( $\beta$ ) values for O<sub>2</sub> and N<sub>2</sub> in H<sub>2</sub>O at STP are 0.028 and 0.014 respectively. There are 55.5 moles of H<sub>2</sub>O in 1L of H<sub>2</sub>O, so the molar ratios of O<sub>2</sub> to H<sub>2</sub>O can be calculated. Note that the concentration of O<sub>2</sub> in liquids is higher when the liquids are in equilibrium with excess O<sub>2</sub> than when they are with excess of air.

The solubility coefficients ( $\beta$ ) and/or ( $I$ ) of some gases in liquids are given in Tables 25–28. Tables of the solubilities of HCl and NH<sub>3</sub> (g/100g of solution) at 760 mm (Table 29) and the boiling points of some useful gases at 760 mm (Table 30) are included.

## MISCELLANEOUS TECHNIQUES

### Freeze-pump-thaw and purging

Volatile contaminants, e.g. traces of low boiling solvent residue or oxygen, in liquid samples or solutions can be very deleterious to the samples on storage. These contaminants can be removed by repeated freeze-pump-thaw cycles. This involves freezing the liquid material under high vacuum in an appropriate vessel (which should be large enough to avoid contaminating the vacuum line with liquid that has bumped) connected to the vacuum line *via* efficient liquid nitrogen traps. The frozen sample is then thawed until it liquefies, kept in this form for some time (*ca* 10–15 minutes), refreezing the sample and the cycle repeated several times without interrupting the vacuum. This procedure applies equally well to solutions, as well as purified liquids, e.g. as a means of removing oxygen from solutions for NMR and other measurements. If the presence of nitrogen, helium or argon, is not a serious contaminant then solutions can be freed from gases, e.g. oxygen, carbon dioxide, and volatile impurities by purging with N<sub>2</sub>, He or Ar at room, or slightly elevated, temperature. The gases used for purging are then removed by freeze-pump-thaw cycles or simply by keeping in a vacuum for several hours. Special NMR tubes with a screw cap thread and a PTFE valve (Wilmad) are convenient for freeze thawing of NMR samples. Alternatively NMR tubes with “J Young” valves (Wilmad) can also be used.

### Vacuum lines, Schlenk and glovebox techniques

Manipulations involving materials sensitive to air or water vapour can be carried out by these procedures. Vacuum line methods make use of quantitative transfers, and **P**(pressure)-**V**(volume)-**T**(temperature) measurements, of gases, and trap-to-trap separations of volatile substances.

It is usually more convenient to work under an inert-gas atmosphere using **Schlenk** type apparatus. The *principle* of Schlenk methods involve all-glass tubes, flasks or vessels which have standard ground-glass joints with one or more side-arms, one of which may have a tap. The system can be purged by evacuating and flushing with an inert gas (usually dry nitrogen, or in some cases, argon or helium), repeating the process until the contaminants, e.g. O<sub>2</sub>, H<sub>2</sub>O or CO<sub>2</sub> in the vapour phases have been diminished to acceptable limits. Many of the reactions using Schlenk equipment require anhydrous conditions and in this case the equipment should be heated in an oven slightly above 100° for 1 to 2 hours (preferably with dry N<sub>2</sub>, He or argon flushing), and allowed to cool to room temperature in the presence of a desiccant. A large range of Schlenk glassware is commercially available. Schlenk equipment in which refluxing of liquids is possible without contact with the atmosphere outside of the apparatus is available commercially. With these, and tailor-made pieces of glassware, inert atmospheres can be maintained during transfer of material, crystallisation, reflux, filtration, and sublimation. Where addition of a solid sample should be made, an L-tube, or a small bulb with a bent neck, with a glass joint is used in which the solid is placed, and can be transferred to the main reaction vessel by simply rotating the tube or bulb. In the case of a liquid, a separating funnel with an equalising tube can be used to allow equilibration of pressure. Also, liquids can be injected, *via* a syringe through “Sure/Seal” caps which can be stretched over, or insert nicely into, the ground joint necks of the main reaction container.

Syringe techniques have been developed for small volumes, while for large volumes or where much manipulation is required, dryboxes (*glove boxes*) or dry chambers should be used. Disposable glove bags (e.g. Atmosbags see Sigma-Aldrich Labware of various dimensions) with two or four hands which can be sealed, purged and inflated with an inert gas are available and are relatively cheap and disposable. They are useful not only for handling moisture-sensitive substances, but also for toxic materials.

## CHEMICAL AND BIOCHEMICAL SOURCES

Apart from wishing to obtain a pure substance there are many reasons for wanting to purify a substance. For example the substance may have been in the store for too long and has deteriorated to a smaller or larger extent and needs to be used. Large quantities may be required, so bulk amounts, less pure but of cheaper grade could serve the purpose if they can be purified readily and cheaply. The cost consideration is very important. Substances that are available commercially can be of varying grades of purity and the purer the grade the higher the price. Biological substances may be only available in crude form, e.g. acetone powders for enzymes. There are a large number of suppliers of substances for chemical, biochemical and for biological requirements and they are continually improving quality, increasing their range and introducing recently developed substances. The following is a website list of the more commonly used suppliers from which almost all the substances and equipment described in this book can be purchased. The list also contains suppliers of laboratory ware as well as of scientific instruments. The list is not exhaustive.

## USEFUL WEBSITES

<http://www.>

chemsupply.com.au [organics, inorganics & equipment]  
sigmaldrich.com/ [organics, inorganics, lifescience materials & equipment]  
wilmad-labglass.com. [glass ware & equipment]  
merck-chemicals.com/ [organics, inorganics & equipment]  
acros.com/ [organics (Acros organics, inorganics & equipment)  
alfa.com/ [organics, inorganics & equipment]  
strem.com/ [general inorganics, metal-organics, catalysts, nanomaterials]  
tci-asiapacific.com/ [Tokyo Chemical Industry, chemicals, lab equipment]  
thermofisher.com/global/en/home.asp [instruments, chemicals, custom products]  
<https://au.vwr.com/app/Home> [VWR International- chemicals/laboratory scientific supplies]  
quantumscientific.com/ [chemical, biochemical & lab equipment]  
gelifesciences.com [GE Healthcare, chemicals, biochemicals & life science products]  
bio-strategy.com [Laboratory technology]  
glschina.com [GL biochemical products]  
invitrogen.com [Invitrogen, life science products]  
lifetechnologies.com [molecular biology products and equipment]  
promega.com [Promega, life science products]  
tocris.com [Tocris Bioscience products, i.e. neurochemicals, biochemical, peptides, DNA]  
novachem.com.au  
[scilabware.com](http://scilabware.com) [plastic labware]  
waters.com [Waters, chromatography materials]  
[biaseparations.com](http://biaseparations.com) [chromatography materials]  
[daicel.co.jp/indexe.html](http://daicel.co.jp/indexe.html) [chromatography materials]  
restekcorp.com [chromatography materials]  
winlab.com.au/ [chromatography materials]  
Fritsch-laser.com [for up to nano particle size and milling]  
retsch-technology.com [for up to nano particle size]  
haverstandard.com [for up to nano particle size]  
[perkinelmer.com](http://perkinelmer.com) [spectral and other instruments]  
agilent.com/chem/atomicsspec/ap  
betterworld.com [books]  
abebooks.com [books]  
booksandcollectibles.com.au [books, collectibles]  
ebay.com.au/ [books, etc]

For high purity inorganic compounds, NIST Traceable inorganic reference standards/calibrants and aqueous standard solutions for ICP, ICP-MS, AA, GFAA and IC visit <[www.exaxol.com](http://www.exaxol.com)>. Note that all the trace metal analyses in the “Inorganic Compounds” section of Chapter 5 are by courtesy of Joe Papa (EXAXOL see Preface).

## TABLES

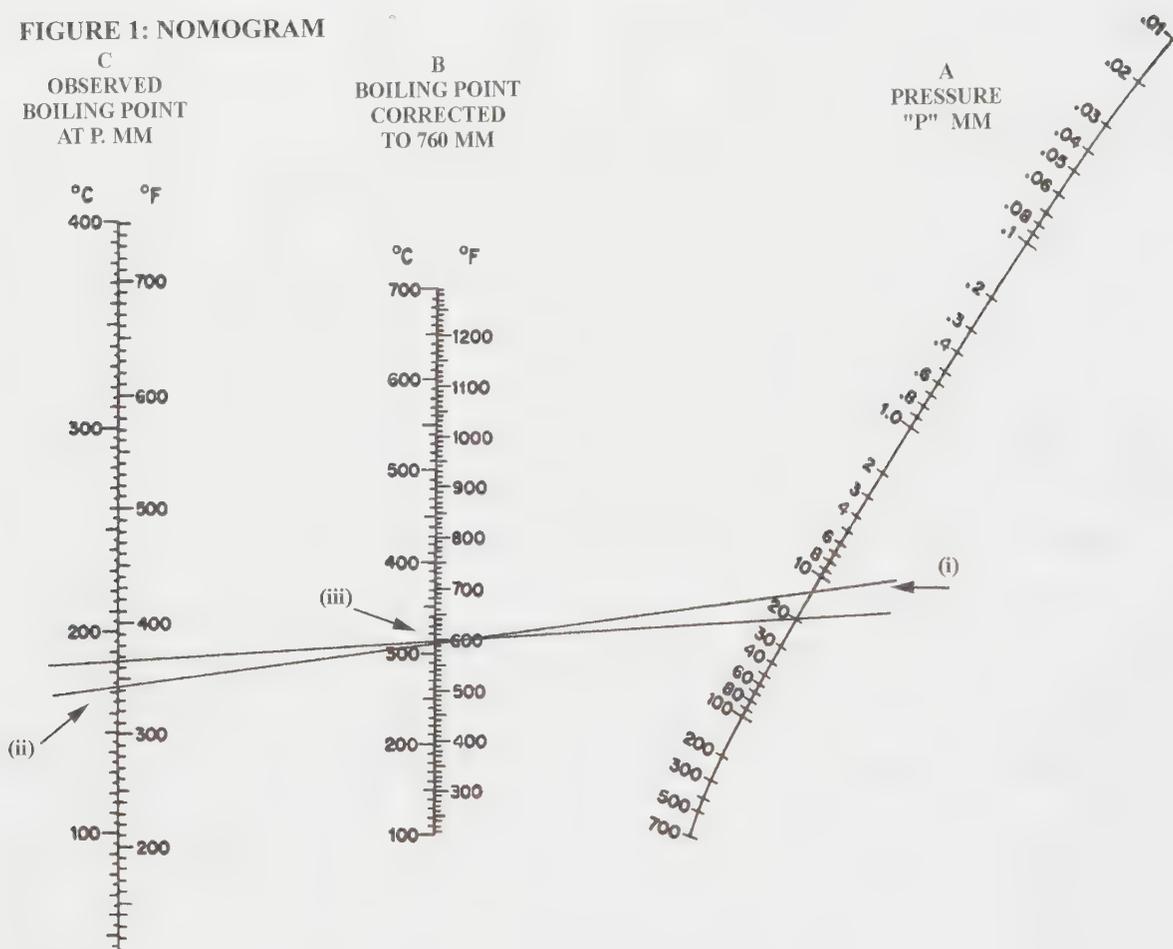
TABLE 1. SOME COMMON IMMISCIBLE OR SLIGHTLY MISCIBLE PAIRS OF SOLVENTS AT AMBIENT TEMPERATURES

---

<b>Acetonitrile</b> with hexane, heptane, iso-octane, cyclohexane.
<b>Benzene</b> with water, brine and aqueous solutions generally.
<b>Butanol</b> with water, brine and aqueous solutions generally.
<b>Carbon tetrachloride</b> with ethanolamine, ethylene glycol, formamide, water or brine.
<b>Chloroform</b> with glycerol, ethylene glycol, water, aqueous solutions generally
<b>Cyclohexane</b> with alcohols, dimethyl formamide, dimethyl sulfoxide, glycerol, pyridine.
<b>Cyclopropyl methyl ether</b> same as ethyl ether but less so.
<b>Dimethyl formamide</b> or <b>dimethyl acetamide</b> with cyclohexane, pentane, petroleum ether, xylene.
<b>Dimethyl sulfoxide</b> with ethyl ether, pentane, petroleum ethers, cyclohexane, xylene.
<b>Ethyl acetate</b> with aqueous solutions generally or petroleum ethers.
<b>Ethyl ether</b> with ethanolamine, dimethyl sulfoxide, ethylene glycol, glycerol, water or aqueous solutions generally.
<b>Ethanol</b> with carbon disulfide, petroleum ethers, cyclohexanes.
<b>Glycerol</b> with benzene, ether, chloroform, carbon tetrachloride, carbon disulfide, petroleum ethers, oils.
<b>Iso-octane</b> with acetonitrile, dimethyl formamide, dimethyl sulfoxide, methanol, water.
<b>Methanol</b> with carbon disulfide, hexane, heptane, cyclohexane or petroleum ethers.
<b>N-Methylpyrrolidone</b> with petroleum ethers, cyclohexanes.
<b>Petroleum ether(s)</b> with aniline, benzyl alcohol, dimethyl formamide, dimethyl sulfoxide, formamide, phenol or water and aqueous solutions generally.
<b>Phenol</b> with petroleum ethers, cyclohexanes.
<b>Pyridine</b> with petroleum ethers, hexanes.
<b>Toluene</b> with water, brine, aqueous solutions generally, glycerol but less so than benzene.
<b>Water</b> with aniline, benzene, benzyl alcohol, carbon disulfide, carbon tetrachloride, chloroform, cyclohexane, cyclohexanol, cyclohexanone, diethyl ether, ethyl acetate, isoamyl alcohol, methyl ethyl ketone, nitromethane, tributyl phosphate or toluene.
<b>Xylene</b> with water, brine, aqueous solutions generally, glycerol, dimethyl formamide, dimethyl sulfoxide.

---

FIGURE 1: NOMOGRAM



How to use Figure 1:

You can use a nomogram to estimate the boiling points of a substance at a particular pressure. For example, the boiling point of 4-methoxybenzenesulfonyl chloride is 173°C/14mm. Thus to find out what the boiling point of this compound will be at 760 mm (atmospheric), draw a point on curve A (pressure) at 14mm (this is shown in (i)). Then draw a point on curve C (observed boiling point) corresponding to 173° (or as close as possible). This is shown in (ii). Using a ruler, find the point of intersection on curve B, drawing a line between points (i) and (ii). This is the point (iii) and is the boiling point of 4-methoxybenzenesulfonyl chloride (i.e. approx. 310°C) at atmospheric pressure. If you want to distil 4-methoxybenzenesulfonyl chloride at 20 mm, then you will need to draw a point on curve A (at 20 mm). Using a ruler, find the point of intersection on curve C drawing through the line intersecting (iii, curve B, i.e. 310°C) and the point in curve A corresponding to 20 mm. You should have a value of 185°C; that is, the boiling point of 4-methoxybenzenesulfonyl chloride is estimated to be at 185°C at 20 mm.

TABLE 2A. PREDICTED EFFECT OF PRESSURE ON BOILING POINT\*

Temperature in degrees Centigrade										
760 mmHg	0	20	40	60	80	100	120	140	160	180
0.1	-111	-99	-87	-75	-63	-51	-39	-27	-15	-4
0.2	-105	-93	-81	-69	-56	-44	-32	-19	-7	5
0.4	-100	-87	-74	-62	-49	-36	-24	-11	2	15
0.6	-96	-83	-70	-57	-44	-32	-19	-6	7	20
0.8	-94	-81	-67	-54	-41	-28	-15	-2	11	24
1.0	-92	-78	-65	-52	-39	-25	-12	1	15	28
2.0	-85	-71	-58	-44	-30	-16	-3	11	25	39
4.0	-78	-64	-49	-35	-21	-7	8	22	36	51
6.0	-74	-59	-44	-30	-15	-1	14	29	43	58
8.0	-70	-56	-41	-26	-11	4	19	34	48	63
10.0	-68	-53	-38	-23	-8	7	22	37	53	68
14.0	-64	-48	-33	-23	-2	13	28	44	59	74
16.0	-61	-45	-29	-14	2	17	33	48	64	79
20.0	-59	-44	-28	-12	3	19	35	50	66	82
30.0	-54	-38	-22	-6	10	26	42	58	74	90
40.0	-50	-34	-17	-1	15	32	48	64	81	97
50.0	-47	-30	-14	3	19	36	52	69	86	102
60.0	-44	-28	-11	6	23	40	56	73	86	107
80.0	-40	-23	-6	11	28	45	62	79	97	114
100.0	-37	-19	-2	15	33	50	67	85	102	119
150.0	-30	-12	6	23	41	59	77	95	112	130
200.0	-25	-7	11	29	47	66	84	102	120	138
300.0	-18	1	19	38	57	75	94	113	131	150
400.0	-13	6	25	44	64	83	102	121	140	159
500.0	-8	11	30	50	69	88	108	127	147	166
600.0	-5	15	34	54	74	93	113	133	152	172
700.0	-2	18	38	58	78	98	118	137	157	177
750.0	0	20	40	60	80	100	120	140	160	180
770.0	0	20	40	60	80	100	120	140	160	180
800.0	1	21	41	61	81	101	122	142	162	182

\* *How to use the Table:* Take as an example a liquid with a boiling point of 80°C at 760 mmHg. The Table gives values of the boiling points of this liquid at pressures from 0.1 to 800 mmHg. Thus at 50 mmHg this liquid has a boiling point of 19°C, and at 2 mmHg its boiling point would be -30°C.

TABLE 2B. PREDICTED EFFECT OF PRESSURE ON BOILING POINT\*

Temperature in degrees Centigrade											
760 mmHg	200	220	240	260	280	300	320	340	360	380	400
<b>0.1</b>	8	20	32	44	56	68	80	92	104	115	127
<b>0.2</b>	17	30	42	54	67	79	91	103	116	128	140
<b>0.4</b>	27	40	53	65	78	91	103	116	129	141	154
<b>0.6</b>	33	40	59	72	85	98	111	124	137	150	163
<b>0.8</b>	38	51	64	77	90	103	116	130	143	156	169
<b>1.0</b>	41	54	68	81	94	108	121	134	147	161	174
<b>2.0</b>	53	66	80	94	108	121	135	149	163	176	190
<b>4.0</b>	65	79	93	108	122	136	151	156	179	193	208
<b>6.0</b>	72	87	102	116	131	146	160	175	189	204	219
<b>8.0</b>	78	93	108	123	137	152	167	182	197	212	227
<b>10.0</b>	83	98	113	128	143	158	173	188	203	218	233
<b>14.0</b>	90	105	120	136	151	166	182	197	212	228	243
<b>18.0</b>	95	111	126	142	157	173	188	204	219	235	251
<b>20.0</b>	97	113	129	144	160	176	191	207	223	238	254
<b>30.0</b>	106	123	139	155	171	187	203	219	235	251	267
<b>40.0</b>	113	130	146	162	179	195	211	228	244	260	277
<b>50.0</b>	119	135	152	168	185	202	218	235	251	268	284
<b>60.0</b>	123	140	157	174	190	207	224	241	257	274	291
<b>80.0</b>	131	148	165	182	199	216	233	250	267	284	301
<b>100.0</b>	137	154	171	189	206	223	241	258	275	293	310
<b>150.0</b>	148	166	184	201	219	237	255	273	290	308	326
<b>200.0</b>	156	174	193	211	229	247	265	283	302	320	338
<b>300.0</b>	169	187	206	225	243	262	281	299	318	337	355
<b>400.0</b>	178	197	216	235	254	273	292	311	330	350	369
<b>500.0</b>	185	205	224	244	263	282	302	321	340	360	379
<b>600.0</b>	192	211	231	251	270	290	310	329	349	368	388
<b>700.0</b>	197	217	237	257	277	296	316	336	356	376	396
<b>750.0</b>	200	220	239	259	279	299	319	339	359	379	399
<b>770.0</b>	200	220	241	261	281	301	321	341	361	381	401
<b>800.0</b>	202	222	242	262	282	302	322	342	362	382	403

\*How to use the Table: Taking as an example a liquid with a boiling point of 340°C at 760 mmHg, the column headed 340°C gives values of the boiling points of this liquid at each value of pressures from 0.1 to 800 mmHg. Thus, at 100 mmHg its boiling point is 258°C, and at 0.8mm Hg its boiling point will be 130°C.

**TABLE 3. HEATING BATHS**

Up to 100°	Water baths
20 to 200°	Glycerol or di- <i>n</i> -butyl phthalate
Up to about 200°	Medicinal paraffin
Up to about 250°	Hard hydrogenated cotton-seed oil (m 40–60°) or a 1:1 mixture of cotton-seed oil and castor oil containing about 1% of hydroquinone.
40 to 250° (to 400° under N <sub>2</sub> )	D.C. 550 silicone fluid
Up to about 260°	A mixture of 85% orthophosphoric acid (4 parts) and metaphosphoric acid (1 part)
Up to 340°	A mixture of 85% orthophosphoric acid (2 parts) and metaphosphoric acid (1 part)
60 to 500°	Fisher bath wax (highly unsaturated)
73 to 350°	Wood's Metal*
250 to 800°	Solder*
350 to 800°	Lead*

\* In using metal baths, the container (usually a metal crucible) should be removed while the metal is still molten.

**TABLE 4. WHATMAN FILTER PAPERS**

Grade No.	1	2	3	4	5	6	113
Particle size retained (in microns)	11	8	5	12	2.4	2.8	28
Filtration speed*(sec/100ml)	40	55	155	20	<300	125	9

**Routine ashless filters**

Grade No.	40	41	42	43	44
Particle size retained (in microns)	7.5	12	3	12	4
Filtration speed* (sec/100ml)	68	19	200	38	125

**Hardened****Hardened ashless**

Grade No.	50	52	54	540	541	542
Particle size retained(in microns)	3	8	20	9	20	3
Filtration speed* (sec/100ml)	250	55	10	55	12	250

**Glass microfilters**

Grade No	GF/A	GF/B	GF/C	GF/D	GF/F
Particle size retained (in microns)	1.6	1.0	1.1	2.2	0.8
Filtration speed (sec/100ml)*	8.3	20.0	8.7	5.5	17.2

\*Filtration speeds are rough estimates of initial flow rates and should be considered on a relative basis.

TABLE 5. MICRO FILTERS\*

Nucleopore (polycarbonate) Filters						
Mean Pore Size (microns)	8.0	2.0	1.0	0.1	0.03	0.015
Av. pores/cm <sup>2</sup>	10 <sup>5</sup>	2×10 <sup>6</sup>	2×10 <sup>7</sup>	3×10 <sup>8</sup>	6×10 <sup>8</sup>	1–6×10 <sup>9</sup>
Water flow rate(ml/min/cm <sup>2</sup> )	2000	2000	300	8	0.03	0.1–0.5
Millipore Filters						
Type	—Cellulose ester—		—Teflon—		—Microweb <sup>#</sup> —	
	MF/SC	MF/VF	LC	LS	WS	WH
Mean Pore Size (microns)	8	0.01	10	5	3	0.45
Water flow rate (ml/min/cm <sup>2</sup> )	850	0.2	170	70	155	55
Gelman Membranes						
Type	—Cellulose ester—		—Copolymer—			
	GA-1	T CM-450	VM-1	DM-800	AN-200	Tuffryn-450
Mean Pore Size (microns)	5	0.45	5	0.8	0.2	0.45
Water flow rate (ml/min/cm <sup>2</sup> )	320	50	700	200	17	50
Sartorius Membrane Filters (SM)						
Application	Gravi-metric	Biological clarification	Sterilisation	Particle count in H <sub>2</sub> O	For acids & bases	
Type No.	11003	11004	11006	11011	12801	
Mean Pore Size (microns)	1.2	0.6	0.45	0.01	8.0	
Water flow rate (ml/min/cm <sup>2</sup> )	300	150	65	0.6	1100	

\* Only a few representative filters are tabulated (available ranges are more extensive). # Reinforced nylon.

**TABLE 6. COMMON SOLVENTS USED IN RECRYSTALLISATION**

Acetic acid (118°)	*Cyclohexane (81°)	*Methanol (64.5°)
*Acetone (56°)	Dichloromethane (41°)	*Methyl ethyl ketone (80°)
Acetylacetone (139°)	*Diethyl ether (34.5°)	Methyl isobutyl ketone (116°)
Acetonitrile (82°C)	Dimethyl formamide (76°/39mm)	Nitrobenzene (210°)
*Benzene (80°)	*Dioxane (101°)	Nitromethane (101°)
Benzyl alcohol (93°/10mm)	*Ethanol (78°)	*Petroleum ether (various)
<i>n</i> -Butanol (118°)	2-Ethoxyethanol (cellosolve 135°)	Pyridine (115.5°)
Butyl acetate (126.5°)	*Ethyl acetate (78°)	Pyridine trihydrate (93°)
<i>n</i> -Butyl ether (142°)	Ethyl benzoate (98°/19mm)	*Tetrahydrofuran (64-66°)
$\gamma$ -Butyrolactone (206°)	Ethylene glycol (68°/4mm)	Toluene (110°)
Carbon tetrachloride (77°)	Formamide (110°/10mm)	Trimethylene glycol (59°/11mm)
Chlorobenzene (132°)	Glycerol (126°/11mm)	Water (100°)
Chloroform (61°)	Isoamyl alcohol (131°)	Xylenes ( <i>o</i> 143-145°, <i>m</i> 138-139°, <i>p</i> 138°)

\*Highly flammable, should be heated or evaporated on steam or electrically heated water baths only (preferably under nitrogen). None of these solvents should be heated over a naked flame.

**TABLE 7. PAIRS OF MISCIBLE SOLVENTS**

**Acetic acid:** with chloroform, ethanol, ethyl acetate, acetonitrile, petroleum ether, or water.

**Acetone:** with benzene, butyl acetate, butyl alcohol, carbon tetrachloride, chloroform, cyclohexane, ethanol, ethyl acetate, methyl acetate, acetonitrile, petroleum ether or water.

**Ammonia:** with ethanol, methanol, pyridine.

**Aniline:** with acetone, benzene, carbon tetrachloride, ethyl ether, *n*-heptane, methanol, acetonitrile or nitrobenzene.

**Benzene:** with acetone, butyl alcohol, carbon tetrachloride, chloroform, cyclohexane, ethanol, acetonitrile, petroleum ether or pyridine.

**Butyl alcohol:** with acetone or ethyl acetate.

**Carbon disulfide:** with petroleum ether.

**Carbon tetrachloride:** with cyclohexane.

**Chloroform:** with acetic acid, acetone, benzene, ethanol, ethyl acetate, hexane, methanol or pyridine.

**Cyclohexane:** with acetone, benzene, carbon tetrachloride, ethanol or diethyl ether.

**Diethyl ether:** with acetone, cyclohexane, ethanol, methanol, methylal (dimethoxymethane), acetonitrile, pentane or petroleum ether.

**Dimethyl formamide:** with benzene, ethanol or ether.

**Dimethyl sulfoxide:** with acetone, benzene, chloroform, ethanol, diethyl ether or water.

**Dioxane:** with benzene, carbon tetrachloride, chloroform, ethanol, diethyl ether, petroleum ether, pyridine or water.

**Ethanol:** with acetic acid, acetone, benzene, chloroform, cyclohexane, dioxane, ethyl ether, pentane, toluene, water or xylene.

**Ethyl acetate:** with acetic acid, acetone, butyl alcohol, chloroform, or methanol.

**Glycerol:** with ethanol, methanol or water.

**Hexane:** with benzene, chloroform or ethanol.

**Methanol:** with chloroform, diethyl ether, glycerol or water.

**Methylal:** with diethyl ether.

**Methyl ethyl ketone:** with acetic acid, benzene, ethanol or methanol.

**Nitrobenzene:** with aniline, methanol or acetonitrile.

**Pentane:** with ethanol or diethyl ether.

**Petroleum ether:** with acetic acid, acetone, benzene, carbon disulfide or diethyl ether.

**Phenol:** with carbon tetrachloride, ethanol, diethyl ether or xylene.

**Pyridine:** with acetone, ammonia, benzene, chloroform, dioxane, petroleum ether, toluene or water.

**Toluene:** with ethanol, diethyl ether or pyridine.

**Water:** with acetic acid, acetone, ethanol, methanol, or pyridine.

**Xylene:** with ethanol or phenol.

TABLE 8. MATERIALS FOR COOLING BATHS

Temp. (°)	Composition	Temp. (°)	Composition
0	Crushed ice	-72	Solid CO <sub>2</sub> with ethanol
-5 to -20	Ice-salt mixtures	77	Solid CO <sub>2</sub> with chloroform or acetone
Up to -20	Ice-MeOH mixtures	-78	Solid CO <sub>2</sub> (powdered; CO <sub>2</sub> snow)
-33	Liquid ammonia	-100	Solid CO <sub>2</sub> with diethyl ether
-40 to -50	Ice (3.5-4 parts) - CaCl <sub>2</sub> ·6H <sub>2</sub> O (5 parts)	-196	liquid nitrogen (see footnote*)

Alternatively, the following liquids can be used, partially frozen, as cryostats, by adding solid CO<sub>2</sub> from time to time to the material in a Dewar-type container and stirring to make a slush:

13	<i>p</i> -Xylene	-55	Diacetone
12	Dioxane	-56	<i>n</i> -Octane
6	Cyclohexane	-60	Di-isopropyl ether
5	Benzene	-73	Trichloroethylene or isopropyl acetate
2	Formamide	-74	<i>o</i> -Cymene or <i>p</i> -cymene
-8.6	Methyl salicylate	-77	Butyl acetate
-9	Hexane-2,5-dione	-79	Isoamyl acetate
-10.5	Ethylene glycol	-83	Propylamine
-11.9	<i>tert</i> -Amyl alcohol	-83.6	Ethyl acetate
-12	Cycloheptane or methyl benzoate	-86	Methyl ethyl ketone
-15	Benzyl alcohol	-89	<i>n</i> -Butanol
-16.3	<i>n</i> -Octanol	-90	Nitroethane
-18	1,2-Dichlorobenzene	-91	Heptane
-22	Tetrachloroethylene	-92	<i>n</i> -Propyl acetate
-22.4	Butyl benzoate	-93	2-Nitropropane or cyclopentane
-22.8	Carbon tetrachloride	-94	Ethyl benzene or hexane
-24.5	Diethyl sulfate	94.6	Acetone
-25	1,3-Dichlorobenzene	-95.1	Toluene
-29	<i>o</i> -Xylene or pentachloroethane	-97	Cumene
-30	Bromobenzene	-98	Methanol or methyl acetate
-32	<i>m</i> -Toluidine	-99	Isobutyl acetate
-32.6	Dipropyl ketone	-104	Cyclohexene
-38	Thiophene	-107	Isooctane
-41	Acetonitrile	-108	1-Nitropropane
-42	Pyridine or diethyl ketone	-116	Ethanol or diethyl ether
-44	Cyclohexyl chloride	-117	Isoamyl alcohol
-45	Chlorobenzene	-126	Methylcyclohexane
-47	<i>m</i> -Xylene	-131	<i>n</i> -Pentane
-50	Ethyl malonate or <i>n</i> -butylamine	-160	Isopentane
-52	Benzyl acetate or diethylcarbitol		

For other organic materials used in low temperature slush-baths with liquid nitrogen see R.E.Rondeau [*J Chem Eng Data* **11** 124 1966]. \*NOTE: Use high quality pure nitrogen; do not use liquid air or liquid nitrogen that has been in contact with air for a long period (due to the dissolution of oxygen in it) as this could EXPLODE in contact with organic matter.

TABLE 9. LIQUIDS FOR STATIONARY PHASES IN GAS CHROMATOGRAPHY\*

Material	Temp. (°)	Retards
Dimethylsulfolane	0–40	Olefins and aromatic hydrocarbons
Di- <i>n</i> -butyl phthalate	0–40	General purposes
Squalane	0–150	Volatile hydrocarbons and polar molecules
Silicone oil or grease	0–250	General purposes
Diglycerol	20–120	Water, alcohols, amines, esters, and aromatic hydrocarbons
Dinonyl phthalate	20–130	General purposes
Polydiethylene glycol succinate	50–200	Aromatic hydrocarbons, alcohols, ketones, esters
Polyethylene glycol	50–200	Water, alcohols, amines, esters and aromatic hydrocarbons
Apiezon grease	50–200	Volatile hydrocarbons and polar molecules
Tricresyl phosphate	50–250	General purposes

\* See Suppliers at the end of section on HPLC and E.F. Barry and R.L. Grob, *Columns for Gas Chromatography*, J. Wiley and Sons NY, 2007, ISBN 9780471740438.

TABLE 10. METHODS OF VISUALISATION OF TLC SPOTS

Reagent	Compound	Preparation	Observations
Iodine	General	Iodine crystals in a closed chamber or spray 1% methanol solution of Iodine	Brown spots which may disappear upon standing
H <sub>2</sub> SO <sub>4</sub>	General	50% solution, followed by heating to 150°C	Black or coloured spots
Molybdate	General	5% (NH <sub>4</sub> ) <sub>6</sub> Mo <sub>7</sub> O <sub>24</sub> + 0.2% Ce(SO <sub>4</sub> ) <sub>2</sub> in 5% H <sub>2</sub> SO <sub>4</sub> , followed by heating to 150°C	Deep blue spots
Vanillin	General	0.5g vanillin, 0.5 ml H <sub>2</sub> SO <sub>4</sub> , 9 ml ethanol	various coloured spots
Ammonia	phenols	Ammonia vapour in a closed chamber	various coloured spots
FeCl <sub>3</sub>	phenols, enols	1% aqueous FeCl <sub>3</sub>	various coloured spots
2,4-DNP	aldehydes, ketones	0.5% 2,4-dinitrophenylhydrazine/2M HCl	red to yellow spots
HCl	aromatic acids and amines	HCl vapour in a closed chamber	various coloured spots
Ninhydrin	amino acids, and amines	0.3% ninhydrin in <i>n</i> -BuOH with 3% AcOH, followed by heating to 125°C/10 min	blue spots
PdCl <sub>2</sub>	S and Se compds	0.5% aq. PdCl <sub>2</sub> + few drops of conc. HCl	red and yellow spots
Anisaldehyde	carbohydrates	0.5 ml anisaldehyde in 0.5 ml conc H <sub>2</sub> SO <sub>4</sub> + 95% EtOH + a few drops of AcOH Heat at 100–110°C for 20–30 minutes	various blue spots

TABLE 11. GRADED ADSORBENTS AND SOLVENTS FOR CHROMATOGRAPHY

Adsorbents (decreasing effectiveness)	Solvents (increasing eluting ability)
Fuller's earth (hydrated aluminosilicate)	Petroleum ether, b. 40–60°.
Magnesium oxide	Petroleum ether, b. 60–80°.
Charcoal	Carbon tetrachloride
Alumina	Cyclohexane
Magnesium trisilicate	Benzene
Silica gel	Diethyl ether
Calcium hydroxide	Chloroform
Magnesium carbonate	Ethyl acetate
Calcium phosphate	Acetone
Calcium carbonate	Ethanol
Sodium carbonate	Methanol
Talc	Pyridine
Inulin	Acetic acid
Sucrose = starch	

TABLE 12. REPRESENTATIVE ION-EXCHANGE RESINS USED IN CHROMATOGRAPHY

<b>Sulfonated polystyrene</b>	<b>Aliphatic amine-type</b>
<b>Strong-acid cation exchanger</b>	<b>weak base anion exchangers</b>
AG 50W-x8	Amberlites IR-45 and IRA-67
Amberlite IR-120	Dowex 3-x4A
Dowex 50W-x8	Permutit E
Duolite 225	Permutit A 240A
Permutit RS	
Permutite C50D	
<b>Carboxylic acid-type</b>	<b>Strong Base, anion exchangers</b>
<b>Weak acid cation exchangers</b>	AG 2x8
Amberlite IRC-50	Amberlite IRA-400
Bio-Rex 70	Dowex 2-x8
Chelex 100	Duolite 113
Duolite 436	Permutit ESB
Permutit C	Permutite 330D
Permutits H and H-70	

TABLE 13. MODIFIED FIBROUS CELLULOSES FOR ION-EXCHANGE CHROMATOGRAPHY

<b>Cation exchange</b>	<b>Anion exchange</b>
CM cellulose (carboxymethyl)	DEAE cellulose (diethylaminoethyl)
CM 22, 23 cellulose	DE 22, 23 cellulose
P cellulose (phosphate)	PAB cellulose ( <i>p</i> -aminobenzyl)
SE cellulose (sulfoethyl)	TEAE cellulose (triethylaminoethyl)
SM cellulose (sulfomethyl)	ECTEOA cellulose

*SE* and *SM* are much stronger acids than *CM*, whereas *P* has two ionisable groups (pK 2–3, 6–7), one of which is stronger, the other weaker, than for *CM* (3.5–4.5). For basic strengths, the sequence is: TEAE » DEAE (pK 8-95) > ECTEOA (pK 5.5-7) > PAB. Their exchange capacities lie in the range 0.3 to 1.0 mg equiv/g.

TABLE 14. BEAD FORM ION-EXCHANGE PACKAGINGS FOR CHROMATOGRAPHY<sup>1</sup>

Cation exchange	Capacity (meq/g)	Anion exchange	Capacity (meq/g)
CM-Sephadex C-25, C-50. <sup>2</sup> (weak acid)	4.5±0.5	DEAE-Sephadex A-25, A-50. <sup>7</sup> (weak base)	3.5±0.5
SP-Sephadex C-25, C-50. <sup>3</sup> (strong acid)	2.3±0.3	QAE-Sephadex A-25, A-50. <sup>8</sup> (strong base)	3.0±0.4
CM-Sepharose CL-6B. <sup>4</sup>	0.12±0.02	DEAE-Sepharose CL-6B. <sup>4</sup>	0.13±0.02
		DEAE-Sephacel. <sup>9</sup>	1.4±0.1
Fractogel EMD, CO <sub>2</sub> <sup>-</sup> (pK ~ 4.5), SO <sub>3</sub> <sup>2-</sup> (pK ~ <1). <sup>5</sup>		Fractogel EMD, DMAE (pK ~9), DEAE (pK ~10.8), TMAE (pK >13). <sup>5</sup>	
CM-32 Cellulose.		DE-32 Cellulose.	
CM-52 Cellulose. <sup>6</sup>		DE-52 Cellulose	

<sup>1</sup>May be sterilised by autoclaving at pH 7 and below 120°. <sup>2</sup>Carboxymethyl. <sup>3</sup>Sulfopropyl. <sup>4</sup>Crosslinked agarose gel, no pre-cycling required, pH range 3-10. <sup>5</sup>Hydrophilic methacrylate polymer with very little volume change on change of pH (equivalent to *Toyopearl*, Sigma), available in superfine 650S, and medium 650M particle sizes. <sup>6</sup>Microgranular, pre-swollen, no pre-cycling required. <sup>7</sup>Diethylaminoethyl. <sup>8</sup>Diethyl(2-hydroxy-propyl)aminoethyl. <sup>9</sup>Bead form cellulose, pH range 2-12, no pre-cycling required. Sephadex and Sepharose from GE Healthcare, Fractogel from Merck, Cellulose from Whatman.

TABLE 15 SELECTED CHIRAL COLUMNS FOR CHROMATOGRAPHY\*

Column name	Chiral ligand	Attributes	Manufacture
CHIRA-chrom-1	R- or S- phenylglycine S-leucine	High capacity & efficiency	HICHROM
CHIRA-chrom-2	dinitrophenyl tartramide	Used in wide pH range	CHROM TECH
CHIRAL-AGP	α 1-acid glycoprotein		
CHIRAL-CBH	cellobiohydrolase		
CHIRAL-HAS	human serum albumin		
CHIRALCEL OD/OJ	cellulose derivative	Particular separations versatile	
CHIRALPAK AD/AS/H	amylose derivative	broad range	DIACEL
CHIRALPAK 1A	immobilised amylose	broad range	
CHIRALPAK 1B	immobilised cellulose	for amino acids and primary amines	
CROWNPAK	18-crown 6 type ether		
CHIROFIL	(18-crown 6)(CO <sub>2</sub> H) <sub>4</sub>	for amino acids and primary amines	RStech

(continued)

TABLE 15 (Continued) SELECTED CHIRAL COLUMNS FOR CHROMATOGRAPHY\*

Column name	Chiral ligand	Attributes	Manufacture
CHIROBIOTIC R	Ristocetin A	Broad selectivity	
CHIROBIOTIC T	Teicoplanin	Broad selectivity	Astech 1
CHIROBIOTIC V	Vancomycin	Broad selectivity	
CYCLOBOND I	$\beta$ -cyclodextrin	form chiral	Astech 1
CYCLOBOND II	$\gamma$ -cyclodextrin	inclusion complexes	
KROMASIL DMB	Acetylated <i>N,N'</i> -diallyl	stable, high capacity	Eka Chemicals
KROMASIL TBB	<i>S</i> -tartardiamide	for large prep work	
NUCLEODEX $\beta$ -OH	$\beta$ -cyclodextrin	reverse phase work	
NUCLEODEX $\alpha$ -, $\beta$ - and $\gamma$ -PM	permethylated $\alpha$ -, $\beta$ - and $\gamma$ -cyclodextrins	reverse phase work reverse phase work	Macherey–Nagel
NUCLEOSIL Chiral-1	<i>S</i> -hydroxyproline-Cu <sup>2+</sup> complex	ligand exchange, e.g. $\alpha$ -amino acids	
RESOLVOSIL BSA-7PX	bovine serum albumen	various applications	
ULTRON ES-OVM	ovomucoid protein	stable phase	Shinwa Chem Industries
ULTRON ES-Pepsin	pepsin protein	stable phase	
DACH-DNB	3,5-dinitrobenzoyl derivs	$\pi$ -electron acceptor/ donor — widely used	
ULMO	3,5-dinitrobenzoyl derivs		
$\alpha$ -Burke 2	3,5-dinitrobenzoyl derivs	$\pi$ -electron acceptor	
$\beta$ -GEM 1	3,5-dinitrobenzoyl derivs	$\pi$ -electron acceptor	
Leucine	3,5-dinitrobenzoyl derivs	$\pi$ -electron acceptor	REGIS
Phenyglycine	$\beta$ -lactamase chiral selector	$\pi$ -electron acceptor	
PIRKLE-1J	<i>S</i> -naphthylleucine	$\pi$ -electron donor	
Naphthylleucine	ligand exchange	good for underivatised	
DAVAKOV	ligand exchange	amino acids	

\* These data were generously provided by Gordon Wingate, Operational Director of Winlab Pty Ltd., POBox 5007, Brendale, Queensland 4500, Australia. Further details about these and other chromatographic columns may be obtained from him, and at < <http://www.winlab.com.au/> >.

TABLE 16. LIQUIDS FOR DRYING PISTOLS

Boiling points (760 mm)	(°)	Boiling points (760 mm)	(°)
Ethyl chloride	12.2	Toluene	110.5
Dichloromethane	39.8	Tetrachloroethylene	121.2
Acetone	56.1	Chlorobenzene	132.0
Chloroform	62.0	<i>m</i> -Xylene	139.3
Methanol	64.5	Isoamyl acetate	142.5
Carbon tetrachloride	76.5	Tetrachloroethane	146.3
Ethanol	78.3	Bromobenzene	155.0
Benzene	79.8	<i>p</i> -Cymene	176.0
Trichloroethylene	86.0	<i>o</i> -Chlorobenzene	180.5
Water	100.0	Tetralin	207.0

**TABLE 17. VAPOUR PRESSURES (mm Hg) OF SATURATED AQUEOUS SOLUTIONS IN EQUILIBRIUM WITH SOLID SALTS**

Salt	Temperature					% Humidity at 20°
	10°	15°	20°	25°	30°	
LiCl.H <sub>2</sub> O			2.6			15
CaBr <sub>2</sub> .6H <sub>2</sub> O	2.1	2.7	3.3	4.0	4.8	19
KOAc			3.5			20
CaCl <sub>2</sub> .6H <sub>2</sub> O	3.5	4.5	5.6	6.9	8.3	20
CrO <sub>3</sub>			6.1			32
Zn(NO <sub>3</sub> ) <sub>2</sub> .6H <sub>2</sub> O			7.4			42
K <sub>2</sub> CO <sub>3</sub> .2H <sub>2</sub> O			7.7	10.7		44
KCNS			8.2			47
Na <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> .2H <sub>2</sub> O			9.1			52
Ca(NO <sub>3</sub> ) <sub>2</sub> .4H <sub>2</sub> O	6.0	7.7	9.6	11.9	14.2	55
Mg(NO <sub>3</sub> ) <sub>2</sub> .6H <sub>2</sub> O			9.8			56
NaBr.2H <sub>2</sub> O	5.8	7.8	10.3	13.5	17.5	58
NaNO <sub>2</sub>			11.6			66
NaCl	6.9	9.6	13.2	17.8	21.4	75
NaOAc			13.3			76
NH <sub>4</sub> Cl			13.8			79
(NH <sub>4</sub> ) <sub>2</sub> SO <sub>4</sub>			14.2			81
KBr			14.7			84
KHSO <sub>4</sub>			15.1			86
KCl			15.1	20.2	27.0	86
K <sub>2</sub> CrO <sub>4</sub>			15.4			88
ZnSO <sub>4</sub> .7H <sub>2</sub> O			15.8			90
NH <sub>4</sub> .H <sub>2</sub> PO <sub>4</sub>			16.3			93
KNO <sub>3</sub>			16.7	22.3	29.8	95
Pb(NO <sub>3</sub> ) <sub>2</sub>			17.2			98
H <sub>2</sub> O	9.21	12.79	17.53	23.76	31.82	100

**TABLE 18. DRYING AGENTS FOR CLASSES OF COMPOUNDS**

Class	Dried with
Acetals	Potassium carbonate.
Acids (organic)	Calcium sulfate, magnesium sulfate, sodium sulfate.
Acyl halides	Magnesium sulfate, sodium sulfate.
Alcohols	Calcium oxide, calcium sulfate, magnesium sulfate, potassium carbonate, followed by magnesium and iodine.
Aldehydes	Calcium sulfate, magnesium sulfate, sodium sulfate.
Alkyl halides	Calcium chloride, calcium sulfate, magnesium sulfate, phosphorus pentoxide, sodium sulfate.
Amines	Barium oxide, calcium oxide, potassium hydroxide, sodium carbonate, sodium hydroxide.
Aryl halides	Calcium chloride, calcium sulfate, magnesium sulfate, phosphorus pentoxide, sodium sulfate.
Esters	Magnesium sulfate, potassium carbonate, sodium sulfate.
Ethers	Calcium chloride, calcium sulfate, magnesium sulfate, sodium, lithium aluminium hydride.
Heterocyclic bases	Magnesium sulfate, potassium carbonate, sodium hydroxide.
Hydrocarbons	Calcium chloride, calcium sulfate, magnesium sulfate, phosphorus pentoxide, sodium (not for olefins).
Ketones	Calcium sulfate, magnesium sulfate, potassium carbonate, sodium sulfate.
Mercaptans	Magnesium sulfate, sodium sulfate.
Nitro compounds and Nitriles	Calcium chloride, magnesium sulfate, sodium sulfate.
Sulfides	Calcium chloride, calcium sulfate.

**TABLE 19.**

**STATIC DRYING FOR SELECTED LIQUIDS (25°C)**

Liquid	Water	Linde Type		Activated Alumina	Silicic Acid Gel
		4 A	5 A		
MeOH	Residual H <sub>2</sub> O %	0.54	0.55	—	0.60
	Wt % absorbed	2.50	1.50	—	—
EtOH	Residual H <sub>2</sub> O %	0.25	0.25	0.45	0.68
	Wt % absorbed	7.00	6.80	1.50	—
1-Butylamine	Residual H <sub>2</sub> O %	1.65	1.31	1.93	2.07
	Wt % absorbed	10.40	18.20	3.40	—
2-Ethylhexylamine	Residual H <sub>2</sub> O %	0.25	0.08	0.43	0.53
	Wt % absorbed	15.10	21.10	6.10	1.70
Diethyl ether	Residual H <sub>2</sub> O %	0.001	0.013	0.16	0.27
	Wt % absorbed	9.50	9.20	6.20	4.30
Amyl acetate	Residual H <sub>2</sub> O %	0.002	—	0.33	0.38
	Wt % absorbed	9.30	—	7.40	1.80

TABLE 20

## FLUOROCHROMES\*

	CAS registry No	$\lambda_{\text{ex}}$ (nm) #	$\lambda_{\text{em}}$ (nm)#	Applications
Acridflavin	[8048-52-0]	463	490	Stains many compounds
Allophycocyanin (APC)		604 (650)	650 (660)	Polypeptide with exceptional Fluorescence for cytometry
3-Aminocoumarin	[1635-31-0]	350	445	Forms fluorescent derivatives <i>via</i> reaction with the 3-NH <sub>2</sub> group
Auramine O (basic yellow 2)	[2465-27-2]	440	530	Fluorescent stain for bacteria and probe for certain enzymes, e.g. alcohol dehydrogenases
4-Bromomethyl-7-methoxycoumarin	[35231-44-8]	360 (340)	410 (421)	Forms fluorescent probes with various compds for TLC/HPLC
Chromomycin A <sub>3</sub>	[7059-24-7]	445	575	Stains DNA
Ethidium bromide	[1239-45-8]	493	620	Stains DNA
Fluorescamine	[38183-12-9]			Reacts with various R-NH <sub>2</sub> to form fluorescent compds
Fluorescein	[2321-07-5]	490	514	Fluorescent nucleus for staining a Variety of compounds
Fluorescein- Isothiocyanate (FITC)	[3326-32-7]	492	518	Microsequencing of peptides & proteins, makes fluorescent Amino acids
Hoechst H 2495 (benzoxanthene yellow)	[72845-94-4]			Fluorochrome for various compds
Hoechst H33342 (bisBenzimide 3HCl)	[23491-52-3]	346	640	Intercalates DNA for cytometry and fluorescence microscopy
R-Phycoerythrin	[11016-17-4]	488	572	Forms fluorescent probes with proteins
PKH 2 and 72 (green)		490	504	A Protein Kinase polypeptide that links to cells causing fluorescence
PKH 26 (red)		557	567	Used for labeling cells. (cf PKH 2 or 72)
Tetramethyl rhodamine Isothiocyanate (TRITC)	[95197-95-8]	529 (492)	596 (518)	Protein fluorochrome, used for immunofluorescence

\*\* Of the many fluorochromes reported in this book only a very small selection is tabulated here. # These wavelengths may vary somewhat depending on the solvent used, and on the pH if measured in aqueous solutions.

**TABLE 21. RESIDUAL SOLVENT SIGNALS OF COMMON NMR SOLVENTS**(Adapted from Gottlieb et al. *J Org Chem* 62 7512 1997, Fulmer et al. *Organometallics* 29, 2176 2010.)

Solvents	<sup>1</sup> H NMR	<sup>13</sup> C NMR
Deuterated chloroform, CDCl <sub>3</sub>	7.26	77.16
Deuterated dichloromethane CD <sub>2</sub> Cl <sub>2</sub>	5.32	53.84
Deuterated benzene, C <sub>6</sub> D <sub>6</sub>	7.16	128.06
Deuterated acetone, (CD <sub>3</sub> ) <sub>2</sub> CO	2.05	29.84; 206.26
Deuterated dimethylsulfoxide, (CD <sub>3</sub> ) <sub>2</sub> SO	2.50	35.52
Deuterated acetonitrile, CD <sub>3</sub> CN	1.94	1.32; 118.26
Deuterated methanol, CD <sub>3</sub> OD	3.31	49.00
Deuterated water, D <sub>2</sub> O	4.79	-

**TABLE 22. <sup>1</sup>H NMR CHEMICAL SHIFTS OF TRACE IMPURITIES OF COMMON SOLVENTS AND REAGENTS**

	CDCl <sub>3</sub>	CD <sub>2</sub> Cl <sub>2</sub>	C <sub>6</sub> D <sub>6</sub>	(CD <sub>3</sub> ) <sub>2</sub> CO	(CD <sub>3</sub> ) <sub>2</sub> SO	CD <sub>3</sub> CN	CD <sub>3</sub> OD	D <sub>2</sub> O
H <sub>2</sub> O	1.56	1.52	0.40	2.84	3.33	2.13	4.87	4.79
Acetone	2.17	2.12	1.55	2.05	2.09	2.09	2.15	2.22
Acetonitrile	2.10	1.97	0.58	2.05	2.07	1.94	2.03	2.06
Chloroform	7.26	7.32	6.15	8.02	8.32	7.58	7.90	-
Dichloromethane	5.30	5.33	4.27	5.63	5.76	5.44	5.49	-
Diethyl ether	1.21 (t), 3.48 (q)	1.15 (t), 3.43 (q)	1.11 (t), 3.26 (q)	1.11 (t), 3.41 (q)	1.09 (t), 3.38 (q)	1.12 (t), 3.42 (q)	1.18 (t), 3.49 (q)	1.17 (t), 3.56 (q)
DMF	8.02, 2.96, 2.88	7.96, 2.91, 2.82	7.63, 2.36, 1.86	7.96, 2.94, 2.78	7.95, 2.89, 2.73	7.92, 2.89, 2.77	7.97, 2.99, 2.86	7.92, 3.01, 2.86
DMSO	2.62		1.68	2.52	2.50	2.50	2.65	2.71
Dioxane	3.71	3.65	3.35	3.59	3.57	3.60	3.66	3.75
Ethanol*	1.25 (t), 3.72 (q), 1.32 (s, OH)	1.19 (t), 3.66 (q), 1.33 (s, OH)	0.96 (t), 3.34 (q), 0.50 (s, OH)	1.12 (t), 3.57 (q), 3.39 (s, OH)	1.06 (t), 3.44 (q), 4.63 (s, OH)	1.12 (t), 3.54 (q), 2.47 (s, OH)	1.19 (t), 3.60 (q)	1.17 (t), 3.65 (q)
Ethyl acetate	2.05 (s), 4.12 (q), 1.26 (t)	2.00 (s), 4.08 (q), 1.23 (t)	1.65 (s), 3.89 (q), 0.92 (t)	1.97 (s), 4.05 (q), 1.20 (t)	1.99 (s), 4.03 (q), 1.17 (t)	1.97 (s), 4.06 (q), 1.20 (t)	2.01 (s), 4.09 (q), 1.24 (t)	2.07 (s), 4.14 (q), 1.24 (t)
Methanol*	3.49 (s), 1.09 (s, OH)	3.42 (s), 1.09 (s, OH)	3.07 (s)	3.31 (s), 3.12 (s, OH)	3.16 (s), 4.01 (s, OH)	3.28 (s), 2.16 (s, OH)	3.31	3.34
pyridine	8.62 (m), 7.29 (m), 7.68 (m)	8.59 (m), 7.28 (m), 7.68 (m)	8.53 (m), 6.66 (m), 6.98 (m)	8.58 (m), 7.35 (m), 7.76 (m)	8.58 (m), 7.39 (m), 7.79 (m)	8.57 (m), 7.33 (m), 7.73 (m)	8.53 (m), 7.44 (m), 7.85 (m)	8.52 (m), 7.45 (m), 7.87 (m)
Silicon grease	0.07 (s)	0.09 (s)	0.29 (s)	0.13 (s)	-0.06 (s)	0.08	0.10	-
THF	3.76 (m), 1.85 (m)	3.69 (m), 1.82 (m)	3.57 (m), 1.40 (m)	3.63 (m), 1.79 (m)	3.60 (m), 1.76 (m)	3.64 (m), 1.80 (m)	3.71 (m), 1.87 (m)	3.74 (m), 1.88 (m)
Toluene	2.36 (s), 7.17 (m), 7.25 (m)	2.34 (s), 7.15 (m), 7.24 (m)	2.11 (s), 7.02 (m), 7.13 (m)	2.32 (s), 7.10-7.20 (m), 7.10-7.20 (m)	2.30 (s), 7.18 (m), 7.25 (m)	2.33 (s), 7.10-7.30 (m), 7.10-7.30 (m)	2.32 (s), 7.16 (m), 7.16 (m)	-
Triethylamine	1.03 (t), 2.53 (q)	0.99 (t), 2.48 (q)	0.96 (t), 2.40 (q)	0.96 (t), 2.45 (q)	0.93 (t), 2.43 (q)	0.96 (t), 2.45 (q)	1.05 (t), 2.58 (q)	0.99 (t), 2.57 (q)
Silicon grease	0.07	0.09	0.29	0.13	-0.06	0.08	0.10	-

Shaded region denotes residual undeuterated solvent. Multiplicities are singlet unless denoted otherwise (s=singlet, t=triplet, m=multiplet). The multiplicities of trace impurities may be different for methanol and ethanol if coupling with the OH proton is observed.

TABLE 23.  $^{13}\text{C}$  NMR CHEMICAL SHIFTS OF TRACE IMPURITIES OF COMMON SOLVENTS AND REAGENTS

	$\text{CDCl}_3$	$\text{CD}_2\text{Cl}_2$	$\text{C}_6\text{D}_6$	$(\text{CD}_3)_2\text{CO}$	$(\text{CD}_3)_2\text{SO}$	$\text{CD}_3\text{CN}$	$\text{CD}_3\text{OD}$	$\text{D}_2\text{O}$
Acetone	30.92	31.00	30.14	30.60	30.56	30.91	30.67	30.89
	207.07	206.78	204.43	205.87	206.31	207.43	209.67	215.94
Acetonitrile	1.89	2.03	0.20	1.12	1.03	1.79	0.85	1.47
	116.43	116.92	116.02	117.60	117.91	118.26	118.06	119.68
Chloroform	77.36	77.99	77.79	79.19	79.16	79.17	79.44	-
Dichloromethane	53.52	54.24	53.46	54.95	54.84	55.32	54.78	-
Diethyl ether	15.20	15.44	15.46	15.78	15.12	15.63	15.46	14.77
	65.91	66.11	65.94	66.12	62.05	66.32	66.88	66.42
DMF	31.45	31.39	30.72	31.03	30.73	31.32	31.61	32.03
	36.50	36.56	35.25	36.15	35.73	36.57	36.89	37.54
	162.62	162.57	162.13	162.79	162.29	163.31	164.73	165.53
DMSO	40.76		40.03	41.23	40.45	41.31	40.45	39.39
Dioxane	67.14	67.47	67.16	67.60	66.36	67.72	68.11	67.19
Ethanol	18.41	18.69	18.72	18.89	18.51	18.80	18.40	17.47
	58.28	58.57	57.86	57.72	56.07	57.96	58.26	58.05
Ethyl acetate	14.19	14.37	14.19	14.50	14.40	14.54	14.49	13.92
	21.04	21.15	20.56	20.83	20.68	21.16	20.88	21.15
	60.49	60.63	60.21	60.56	59.74	60.98	61.50	62.32
	171.36	171.24	170.44	170.96	170.31	171.68	172.89	175.26
Methanol	50.41	50.45	49.97	49.77	48.59	49.90	49.86	49.50
Pyridine	123.75	124.06	123.58	124.57	123.84	127.76	125.53	125.12
	135.96	136.16	135.28	136.56	136.05	136.89	138.35	138.27
	149.90	150.27	150.27	150.67	149.58	150.76	150.05	149.18
THF	25.62	25.98	25.72	26.15	25.14	26.27	26.48	25.67
	67.97	68.16	67.80	68.07	67.03	68.33	68.83	68.68
Toluene	21.46	21.53	21.10	21.46	20.99	21.50	21.50	
	125.33	125.62	125.68	126.12	125.29	126.28	126.29	
	128.26	128.54	128.56	129.03	128.18	129.23	129.20	
	129.07	129.35	129.33	129.76	128.88	129.94	129.91	
	137.89	138.36	137.91	138.48	137.35	138.90	138.85	
Triethylamine	11.61	12.12	12.35	12.49	11.74	12.38	11.09	9.07
	46.25	46.75	46.77	47.07	45.74	47.10	46.96	47.19
Silicon grease	1.19	1.22	1.38	1.40			2.10	

TABLE 24. AQUEOUS BUFFERS

Approx. pH	Composition
0	2N sulfuric acid or N hydrochloric acid
1	0.1N hydrochloric acid or 0.18N sulfuric acid
2	<b>Either</b> 0.01N hydrochloric acid or 0.013N sulfuric acid <b>Or</b> 50 ml of 0.1M glycine (also 0.1M NaCl) + 50 ml of 0.1N hydrochloric acid
3	<b>Either</b> 20 ml of the 0.2M Na <sub>2</sub> HPO <sub>4</sub> + 80 ml of 0.1M citric acid <b>Or</b> 50 ml of 0.1M glycine + 22.8 ml of 0.1N hydrochloric acid in 100 ml
4	<b>Either</b> 38.5 ml of 0.2M Na <sub>2</sub> HPO <sub>4</sub> + 61.5 ml of 0.1M citric acid <b>Or</b> 18 ml of 0.2M NaOAc + 82 ml of 0.2M acetic acid
5	<b>Either</b> 70 ml of 0.2M NaOAc + 30 ml of 0.2M acetic acid <b>Or</b> 51.5 ml of 0.2M Na <sub>2</sub> HPO <sub>4</sub> + 48.5 ml of 0.1M citric acid
6	63 ml of 0.2M Na <sub>2</sub> HPO <sub>4</sub> + 37 ml of 0.1M citric acid
7	82 ml of M Na <sub>2</sub> HPO <sub>4</sub> + 18 ml of 0.1M citric acid
8	<b>Either</b> 50 ml of 0.1M Tris buffer + 29 ml of 0.1N hydrochloric acid, in 100 ml <b>Or</b> 30 ml of 0.05M borax + 70 ml of 0.2M boric acid
9	80 ml of 0.05M borax + 20 ml of 0.2M boric acid
10	<b>Either</b> 25 ml of 0.05M borax + 43 ml of 0.1N NaOH, in 100 ml <b>Or</b> 50 ml of 0.1M glycine + 32 ml of 0.1N NaOH, in 100 ml
11	50 ml of 0.15M Na <sub>2</sub> HPO <sub>4</sub> + 15 ml of 0.1N NaOH
12	50 mL of 0.15M Na <sub>2</sub> HPO <sub>4</sub> + 75 ml of 0.1N NaOH
13	0.1N NaOH or KOH
14	N NaOH or KOH

*These buffers are suitable for use in obtaining ultraviolet spectra. Alternatively, for a set of accurate buffers of low, but constant, ionic strength ( $I = 0.01$ ) covering a pH range 2.2 to 11.6 at 20°, see Perrin Aust J Chem 16 572 1963. "In 100 ml" means that the solution is made up to 100 ml with pure water.*

TABLE 25. SOLUBILITY COEFFICIENTS OF AIR AT 1atm IN WATER

T °C	( $\beta$ ) <sub>air</sub>	( $l$ ) <sub>air</sub>	% O <sub>2</sub> (by vol)	( $l$ ) O <sub>2</sub>	Conc of O <sub>2</sub> in H <sub>2</sub> O
0	0.0292	0.0292	34.91	0.0102	0.455mM
20	0.0187	0.0199	34.03	0.0068	0.282mM
25	0.0171	0.0187	33.82	0.0063	0.258mM
30	0.0156	0.0173	33.62	0.0060	0.237mM

( $\beta$ ) is the Bunsen coeff. in ml gas/ml H<sub>2</sub>O at STP. ( $l$ ) is the Ostwald coeff. in ml gas/ml H<sub>2</sub>O at stated temp. Adapted from W.F. Linke, *A. Seidell's Solubilities of Inorganic and Metal-organic Compounds*, American Chemical Society, 1965.

TABLE 26. SOLUBILITY COEFFICIENTS OF O<sub>2</sub> AT 1atm IN WATER

T °C	( $\beta$ )	( $l$ )	Conc of O <sub>2</sub> in H <sub>2</sub> O
0	0.047	0.047	2.10mM
20	0.031	0.033	1.37mM
25	0.028	0.0306	1.25mM
30	0.026	0.0289	1.16mM
50	0.021	0.0248	0.94mM

( $\beta$ ) is the Bunsen coeff. in ml gas/ml H<sub>2</sub>O at STP. ( $l$ ) is the Ostwald coeff. in ml gas/ml H<sub>2</sub>O at stated temp. Adapted from W.F. Linke, *A. Seidell's Solubilities of Inorganic and Metal-organic Compounds*, American Chemical Society, 1965.

TABLE 27. BUNSEN COEFFICIENTS ( $\beta$ ) OF GASES AT 1atm IN ORGANIC SOLVENTS AT 20°C

Solvent	H <sub>2</sub>	He	N <sub>2</sub>	O <sub>2</sub>	CO	CO <sub>2</sub>
H <sub>2</sub> O	0.017	0.009	0.015	0.031	0.025	0.88
CS <sub>2</sub>	0.031	—	0.049	—	0.076	0.83
CHCl <sub>3</sub>	—	—	0.120	0.205	0.177	0.345
EtOH	0.080	0.028	0.130	0.143	0.177	3.0
Me <sub>2</sub> CO	0.065	0.030	0.129	0.207	0.198	6.5
Et <sub>2</sub> O	0.12	—	0.24	0.415	0.38	5.0
C <sub>6</sub> H <sub>6</sub>	0.066	0.018	0.104	0.163	0.153	—

( $\beta$ ) is the Bunsen coeff. in ml gas/ml H<sub>2</sub>O at STP.

Adapted from S. Glasstone, *Textbook of Physical Chemistry*, Macmillan & Co Ltd, London, 1951.

**TABLE 28. OSTWALD COEFFICIENTS (*l*)/L OF O<sub>2</sub> AT 1atm IN AQUEOUS SOLUTIONS AT 25°C**

	( <i>l</i> )/L		( <i>l</i> )
0.125N NH <sub>4</sub> Cl	2,31	0.5M HCl	0.0296
0.25N NH <sub>4</sub> Cl	1.60	1.0M HCl	0.0287
1.0N NH <sub>4</sub> Cl	0.07	2.0M HCl	0.0267
0.125N NaCl	5.52	0.25M H <sub>2</sub> SO <sub>4</sub>	0.0288
0.025N NaCl	5,30	0.5M H <sub>2</sub> SO <sub>4</sub>	0.0275
0.50N NaCl	4.92	1.0M H <sub>2</sub> SO <sub>4</sub>	0.0251
1.0N NaCl	4.20	1.5M H <sub>2</sub> SO <sub>4</sub>	0.0229
2.0N NaCl	3.05	2.0M H <sub>2</sub> SO <sub>4</sub>	0.0209
4.0N NaCl	1.62	2.5M H <sub>2</sub> SO <sub>4</sub>	0.0194
0.125N NaBr	5.65	0.5M HNO <sub>3</sub>	0.0302
0.5N NaBr	5.15	1.0M HNO <sub>3</sub>	0.0295
1.0N NaBr	4.47	2.0M HNO <sub>3</sub>	0.0284
6.0N NaBr	1.28		
0.125N KCl	5.52	0.5M NaOH	0.0250
0.25N KCl	5.30	1.0M NaOH	0.0204
1.0N KCl	4.26	2.0M NaOH	0.0133
4.0N KCl	1.17		
0.125N K <sub>2</sub> SO <sub>4</sub>	5.11	0.5M KOH	0.0252
0.25N K <sub>2</sub> SO <sub>4</sub>	4.66	1.0M KOH	0.0206
0.05N K <sub>2</sub> SO <sub>4</sub>	3.89		
0.125N BaCl <sub>2</sub>	5.40	0.125N Sucrose	0.00540
0.25N BaCl <sub>2</sub>	5.04	0.5N Sucrose	0.00438
1.0N BaCl <sub>2</sub>	3.10	1.0N Sucrose	0.00320
		2.0N Sucrose	0.00184

The Ostwald coefficient (*l*)/L is for ml of gas in 1L of solution, and (*l*) is for ml of gas in 1ml of solution. Adapted from W.F. Linke, *A. Seidell's Solubilities of Inorganic and Metal-organic Compounds*, American Chemical Society, 1965.

TABLE 29. SOLUBILITIES OF HCl AND NH<sub>3</sub> AT 760mm (g/100g OF SOLUTION)

Gas	Temperature °C	MeOH	EtOH	Et <sub>2</sub> O
Hydrogen Chloride*	-10	54.6	—	37.5 (-9.2°)
	0	51.3	45.4	35.6
	20	47.0 (18°)	41.0	24.9
	30	43.0	38.1	19.47
Ammonia	15	21.6 (27.6g/100g MeOH)	13.2 (9.2g/100ml soln)	—
	25	16.5 (19.8g/100g MeOH)	10.0 (6.0g/100ml soln)	—

\* Saturated EtOH with HCl is ~ 5.7M at 25°C, i.e. 21.5g/100ml of solution.

TABLE 30. BOILING POINTS OF SOME USEFUL GASES AT 760 mm

Argon	-185.6°	Krypton	-152.3°
Carbon dioxide (sublimes)	-78.5°	Methane	-164.0°
Carbon monoxide	-191.3°	Neon	-246.0°
Ethane	-88.6°	Nitrogen	-209.9°
Helium	-268.6°	Nitrous oxide	-88.5°
Hydrogen	-252.6°	Nitric oxide	-195.8°
		Oxygen	-182.96°

TABLE 31. PREFIXES FOR QUANTITIES

<b>Fractional</b>	deci (d)	centi (c)	milli (m)	micro (μ)	nano (n)	pico (p)	femto (f)	atto (atto)
	=	=	=	=	=	=	=	=
	10 <sup>-1</sup>	10 <sup>-2</sup>	10 <sup>-3</sup>	10 <sup>-6</sup>	10 <sup>-9</sup>	10 <sup>-12</sup>	10 <sup>-15</sup>	10 <sup>-18</sup>
<b>Multiple</b>	deca (d)	hecto (h)	kilo (k)	mega (M)	giga (G)	tera (T)	penta (P)	eka (E)
	=	=	=	=	=	=	=	=
	10 <sup>1</sup>	10 <sup>2</sup>	10 <sup>3</sup>	10 <sup>6</sup>	10 <sup>9</sup>	10 <sup>12</sup>	10 <sup>15</sup>	10 <sup>18</sup>

## BIBLIOGRAPHY

The following books and reviews provide fuller details of the topics indicated in this chapter. [For earlier bibliographies see *Purification of Laboratory Chemicals*, 4<sup>th</sup>, 5<sup>th</sup>, and 6<sup>th</sup> Edns, ISBN 0750628391 (1996, hardback) and 0750637617 (1997, paperback), 0750675713 (2003), and 9781856175678 (2009). See Bibliography in Chapter 2, and visit the website of DA Books <www.dadirect.com>, Elsevier <www.elsevierdirect.com> and Wiley <www.wiley.com> to update this bibliography. For books that are out of print see <www.Abebooks.com>.

### AFFINITY CHROMATOGRAPHY

P. Bailon, G.K. Ehrlich, W. Fung and W. Berthold (Eds), *Affinity Chromatography: Methods and Protocols*, Humana Press, Totowa, 2000. ISBN 0896036944.

D.S. Hage, *Handbook of Affinity Chromatography*, Marcel Dekker Inc, 2005. ISBN 9780824740573.

P. Matejtschuk (Ed.), *Affinity Separations: A Practical Approach*, Oxford University Press, Oxford, 1997. ISBN 019963551X, 0199635501.

T.M. Phillips and B.F. Dickens, *Affinity and Immunoaffinity Purification Techniques*, Eaton Publishing, Natick MA, 2000. ISBN 1881299228.

M. Zachariou, *Affinity Chromatography: Methods and Protocols, 2nd Edn*, Humana, 2007. ISBN 9781588296597.

### CHIRAL CHROMATOGRAPHY

H.Y. Aboul-Enein, *Chiral Separations by Liquid Chromatography and Related Technologies*, Marcel Dekker Inc, 2003. ISBN 9780824740146.

S. Ahuja (Ed.), *Chiral Separations: Applications and Technology*, ACS, Washington DC, 1997. ISBN 0841234078.

S. Allenmark, *Chromatographic Enantioseparations (Methods and Applications)*, 2nd Edn, Eliss Horwood Publ. New York, 1991. ISBN 0131329782.

T.E. Beesley and R.P.W. Scott, *Chiral Chromatography*, J.Wiley & Son, Chichester, 1999. ISBN 0471974277.

D. Stevenson and I.D. Wilson (Eds), *Recent Advances in Chiral Separations* (Chromatographic Society Symposium Series), Plenum Pub Corp., New York, 1991. ISBN 0306438364.

G. Subramanian (Ed.), *Chiral Separation Techniques: A Practical Approach*, VCH Verlagsgesellschaft, Weinheim, 2000. ISBN 3527298754.

### CHROMATOGRAPHY

R.E. Ardrey, *Liquid Chromatography—Mass Spectrometry: An Introduction*, Wiley & Sons Inc, 2003. ISBN 9780471498018.

K. Blau and J.M. Halket (Eds), *Handbook of Derivatives for Chromatography*, J. Wiley & Sons, New York, 1994. ISBN 047192699X.

P.R. Brown et al. (Eds), *Advances in Chromatography*, Marcel Dekker, New York, Vol 34 1994, and vol 35 1995. ISBN 9780824790875 and 9780824793616 respectively.

- P.R. Brown et al. (Eds), *Advances in Chromatography*, Marcel Dekker, New York, Vols 1-46 1955-2008.
- J. Cazes, *Encyclopedia of Chromatography*, Marcel Dekker, New York, 2001. ISBN 0824705114.
- S.A. Cohen and M.R. Schure (Eds), *Multidimensional Liquid Chromatography: Theory and Applications in Industrial Chemistry and Life Sciences*, J. Wiley & Sons Inc, 2008. ISBN 9780471738473.
- E. Heftmann (Ed.), *Chromatography*, 5th Edn, Elsevier Science, Amsterdam, 1992. ISBN 0444882375.
- E. Heftmann, *Chromatography: Fundamentals and Applications of Chromatography and Related Differential Methods*, Elsevier Science, 2004. ISBN 9780444511072.
- K. Hostettmann, A. Marston and M. Hostettmann, *Preparative Chromatography Techniques: Applications in Natural Product Isolation*, 2nd Edn, Springer-Verlag, Berlin, 1998. ISBN 3540624597.
- P. Miller (Ed.), *High Resolution Chromatography: A Practical Approach*, Oxford University Press, Oxford, 1999. ISBN 0199636486.
- J.M. Miller, *Chromatography: Concepts and Contrasts*, J. Wiley & Sons Inc, 2004. ISBN 9780471472070.
- C.F. Poole and S.K. Poole, *Chromatography Today*, Elsevier Science, Amsterdam, 4th reprint 1997. ISBN 0444891617.
- P.C. Sadek, *Illustrated Pocket Dictionary of Chromatography*, NetLibrary, 2004. eISBN 114739.

## CRYSTALLISATION

- T.M. Bergfors (Ed.), *Protein Crystallisation*, International University Line, 1999. ISBN 0963681753.
- G.D. Botsaris, *Separation & Purification by Crystallization*, American Chemical Society, 1997. ISBN 9780841235137.
- A. Ducruix and R. Giege (Eds), *Crystallisation of Nucleic Acids and Proteins*, IRL Press, Oxford, 1992. ISBN 0199632456. 2nd Edn, Oxford University Press, Oxford, 1999. ISBN 0199636788.
- J. Garside, *From Molecules to Crystallizers; An Introduction to Crystallization*, Oxford University Press, 2001. ISBN 9780198504894.
- A. Mersmann (Ed.), *Crystallisation Technology Handbook*, Marcel Dekker, New York, 1994. ISBN 0824792335.
- H. Michel (Ed.), *Crystallisation of Membrane Proteins*, CRC Press, Boca Raton, 1991. ISBN 0849348161.
- J.M. Mullin, *Crystallisation*, 4th Edn, Butterworth-Heinemann, Oxford, 2001. ISBN 0750648333.
- A.S. Myerson and K. Toyokura, *Crystallisation as a Separation Process*, ACS, Washington DC, 1990. ISBN 0841218641.
- J. Nyvit and J. Ulrich, *Admixtures in Crystallisation*, Wiley-VCH, New York 1995. ISBN 3527287396.

## DISTILLATION

- B. Ding, L. Zhongwei and C. Zhigang, *Special Distillation Processes*, Elsevier Science, 2005. ISBN 9780444516480.

W.L.Luyben, *Distillation: Design and Control Using Aspen Simulation*, J.Wiley and Sons, New York, 2006. ISBN 9780471778882.

W.L. Luyben and C-C. Yu, *Reactive Distillation: Design and Control*, J.Wiley and Sons, New York, 2008. ISBN 9780470226124.

I.M. Mutjaba, *Batch Distillation: Design and Operation*, World Scientific Publ Co, 2004. ISBN 9781860944376.

F. Petlyuk, *Distillation: Theory and Its Application to Optimal Design of Separation Units*, Cambridge University Press UK, 2004. ISBN 9780521820929.

J. Stichlmair and J.R. Fair, *Distillation: Principles and Practices*, J. Wiley and Sons, New York, 1998. ISBN 0471252417.

K. Sundmacher and A. Kienle, *Reactive Distillation: Status and Future Directions*, VCH-Wiley Publishing, 2002. ISBN 9783527305797.

## DRYING

G. Broughton (Chapter VI), in A. Weissberger's (Ed.) *Techniques of Organic Chemistry*, Vol III, pt I, 2nd Edn, Interscience, New York, 1956.

J.F. Coetzee, *Recommended Methods for the Purification of Solvents and Tests for Impurities*, Pergamon Press, Oxford, 1982. ISBN 0080223702.

H. Costantino and M. Pikal, *Lyophilisation of Biopharmaceuticals*, Springer, 2006. ISBN 9780971176768.

G.W. Oetjen, *Freeze-drying*, VCH-Wiley Publishing, 1999. ISBN 0783527295715.

J.A. Riddick and W.B. Bunger, *Organic Solvents: Physical Properties and Methods of Purification, Techniques of Chemistry*, Vol II, Wiley-Interscience, New York, 1970.

## ELECTROPHORESIS

B. Chankvetaze, *Capillary Electrophoresis and Chiral Analysis*, J. Wiley & Sons Inc, 2000, ISBN 9780470865521.

F. Foret, L. Krivankova and K.P. Boce, *Capillary Zone Electrophoresis*, VCH-Wiley Publishing, 1993, ISBN 9783527300198.

A.L. Lagu and M.A. Strega, *Capillary Electrophoresis of Proteins and Peptides*, Marcel Dekker Inc, 2004, ISBN 9781588290175 and 9781592597987.

M.L. Marina, *Analysis and Detection by Capillary Electrophoresis*, Elsevier Science, 2005, ISBN 9780444517180.

K.R. Mitchelson and J. Cheng *Capillary Electrophoresis of Nucleic Acids-Introduction*, Humana, 2004, ISBN 9781592590551.

P. Thibault and S. Honda *Capillary Electrophoresis of Carbohydrates*, Humana, 2003, ISBN 9781592592944.

A.E. Van and E. Michotte, *Chiral Separations by Capillary Electrophoresis*, Taylor & Francis Ltd, 2009, ISBN 9781420069334.

**GAS CHROMATOGRAPHY**

E.F. Barry and R.L. Grob *Columns for Gas Chromatography*, J. Wiley & Sons, New York, 2007, ISBN 9780471740438.

P.J. Baugh (Ed.), *Gas Chromatography: A Practical Approach*, IRL Press, Oxford, 1994. ISBN 0199632715.

R.E. Clement (Ed.), *Gas Chromatography: Biochemical, Biomedical and Clinical Applications*, J. Wiley & Sons, New York, 1990. ISBN 0471010480.

I.A. Fowles, *Gas Chromatography: Analytical Chemistry by Open Learning*, 2nd Edn, J. Wiley & Sons, Chichester, 1995. ISBN 0471954683.

H.H. Hill and D.G. McMinn (Eds), *Detectors for Capillary Chromatography*, J. Wiley & Sons, New York, 1992. ISBN 0471506451.

H-J. Hubschmann, *Handbook of GC/MS: Fundamentals and Applications*, VCH-Wiley Publishing, 2008. ISBN 9783527314270.

W. Jennings, E. Mittlefehldt and P. Stremple, *Analytical Gas Chromatography*, 2nd Edn, Academic Press, San Diego, 1997. ISBN 012384357X.

F.W. Karasek and R.E. Clement, *Basic Gas-Chromatography-Mass Spectrometry: Principles and Techniques*, Elsevier, Amsterdam, 1988. ISBN 0444427600.

M. McMaster and C. McMaster, *GC/MS: A Practical User's Guide*, J. Wiley & Sons, New York, 2008. ISBN 9780470101636.

H.M. McNair and J.M. Miller, *Basic Gas Chromatography: Techniques in Analytical Chemistry*, J. Wiley & Sons, New York, 1997. ISBN 047117260X.

W.M.A. Niessen, *Current Practice of Gas Chromatography*, Marcel Dekker Inc, 2001. ISBN 9780824704735.

D. Rood, *A Practical Guide to the Care, Maintenance, and Troubleshooting of Capillary Gas Chromatographic Systems*, 3rd Edn, Wiley-VCH, New York; 1999. ISBN 3527297502.

R.D. Sacks, J-M.D. Dimandja and D.G. Patterson, *High Speed Gas Chromatography*, J. Wiley & Sons, New York, 2001. ISBN 0471415561.

P. Sandra (Ed.), *Sample Introduction in Capillary Gas Chromatography*, Vol 1, Chromatographic Method Series, Huethig, Basel, 1985.

**GEL FILTRATION**

P.L. Dubin, *Aqueous Size-Exclusion Chromatography*, Elsevier, Amsterdam, 1988. ISBN 0444429573.

L. Fischer, *Gel Filtration*, 2nd Edn, Elsevier, North-Holland, Amsterdam, 1980. ISBN 0444802231.

M. Potschka and P.L. Dubin (Eds), *Strategies in Size-exclusion Chromatography*, ACS, Washington DC, 1996. ISBN 0841234140.

C-S. Wu (Ed.), *Handbook of Size Exclusion Chromatography*, Marcel Dekker, New York, 1995. ISBN 0824792882.

C-S. Wu (Ed.), *Column Handbook for Size Exclusion Chromatography*, Academic Press, New York, 1998. ISBN 0127655557.

W.W. Yau, D.D. By and J.J. Kirkland, *Modern Size Exclusion Liquid Chromatography: Practice of Gel Permeation and Gel Filtration Chromatography*, J. Wiley and Sons, New York, 1979. ISBN 0471033871.

### HIGH PERFORMANCE LIQUID CHROMATOGRAPHY

K.D. Bartle and P. Meyers, *Capillary Electrochromatography*, Royal Society of Chemistry Chromatography Monographs, London, 2000. ISBN 0854045309.

A. Berthold (Ed.), *Countercurrent Chromatography (CCC). The Support-Free Liquid Stationary Phase*, Elsevier, Amsterdam, 2002. ISBN 044450737X.

B.A. Bidlingmeyer, *Preparative Liquid Chromatography*, 2nd Reprint, Elsevier, Amsterdam, 1991. ISBN 0444428321.

R. Freitag, *Modern Advances in Chromatography*, Springer-Verlag, New York, 2002. ISBN 3540430423. [Electronic version available at: <http://link.springer.de/series/abe/>].

T. Hanai and R.M. Smith (Eds), *HPLC: A Practical Guide*, Royal Society of Chemistry, Thomas Graham House, Cambridge, U.K., 1999. ISBN 0854045155.

S. Kromidas, *HPLC Made to Measure: A Practical Handbook for Optimization*, VCH-Wiley Publishing, 2006. ISBN 9783527313778.

W.J. Lough and I.W. Wainer (Eds), *High Performance Liquid Chromatography: Fundamental Principles and Practice*, Blackie Academic and Professional, London, 1996. ISBN 0751400769.

M.C. McMaster, *LC-MS: A Practical User's Guide*, J. Wiley & Sons Inc, 2005. ISBN 9780471655312.

U.D. Neue, *HPLC Columns: Theory, Technology, and Practice*, Wiley-VCH, New York, 1997. ISBN 0471190373.

V. Meyer, *Practical High-Performance Liquid Chromatography*, 2nd Edn., J. Wiley & Sons, New York, 1994. ISBN 0471941328, 0471941298.

R.W.A. Oliver (Ed.), *HPLC of Macromolecules*, 2nd Edn, IRL Press, Oxford, 1998. ISBN 0199635714, and 0199635706.

D. Parriott, *A Practical Guide to HPLC Detection*, Academic Press, San Diego, 1997. ISBN 0125456808.

P.C. Sadek, *Troubleshooting HPLC Systems: A Bench Manual*, Wiley-Interscience, New York, 1999. ISBN 0471178349.

P.C. Sadek, *The HPLC Solvent Guide*, J. Wiley & Sons, New York, 1996. ISBN 0471118559.

L.R. Snyder, J. Kirkland and J. Glajch, *Practical HPLC Method Development*, 2nd Edn., Wiley-Interscience, New York, 1997. ISBN 047100703X.

J. Swadesh (Ed.), *HPLC: Practical and Industrial Applications*, CRC Press, Boca Raton, 1997. ISBN 0849326826.

G. Szepesi, *How to Use Reverse Phase HPLC*, VCH, New York, 1992. ISBN 0895737663.

### IONIC EQUILIBRIA

G. Kortüm, W. Vogel and K. Andrussow, *Dissociation Constants of Organic Acids in Aqueous Solution*, Butterworths, London, 1961.

D.D. Perrin, *Dissociation Constants of Inorganic Acids and Bases in Aqueous Solution*, Butterworths, London, 1969. ISBN 408-70015-7.

D.D. Perrin, *Dissociation Constants of Organic Bases in Aqueous Solution*, Butterworths, London, 1965 suggested additional U.D.C number 543-831-36(083); and Supplement 1972. ISBN 0-408-70408 X.

D.D. Perrin and B. Dempsey, *Buffers for pH and Metal Ion Control*, Chapman and Hall, London, 1974.

W. Sabnis, *Handbook of Acid-Base Indicators*, CRC Press, Boca Raton, Florida, 2007. ISBN 9780849382185.

E.P. Serjeant and B. Dempsey, *Ionisation Constants of Organic Acids in Aqueous Solution*, Pergamon Press, Oxford, 1979. ISBN 0-08-022339-7.

## ION EXCHANGE

A. Dyer, M.J. Hudson and P.A. Williams (Eds), *Ion Exchange Processes: Advances and Applications*, Royal Society of Chemistry, London, 1993. ISBN 0851864457.

J.S. Fritz and D.T. Gjerde, *Ion Chromatography*, 3rd Edn, VCH Verlagsgesellschaft Mbh, Weinheim, 2000. ISBN 3527299149.

J.S. Fritz and D.T. Guide, *Ion Chromatography, 4th Edn, Completely Revised and Enlarged Edition*, J Wiley & Sons Inc, 2008. ISBN 978352732053.

D.T. Gjerde and J.S. Fritz, *Ion Chromatography*, 2nd Edn, J. Wiley and Sons, New York, 1998. ISBN 3527296964.

P.R. Haddad and P.E. Jackson, *Ion Chromatography: Principles and Applications*, Elsevier, New York, 1990. ISBN 0444882324.

F.G. Helfferich, *Ion Exchange*, Dover Publications, New York, 1989. ISBN 0486687848.

J. Korkisch, *Concise Handbook of Ion Exchange Resins in Analytical Chemistry*, CRC Press, Boca Raton, 2001. ISBN 9780849386558.

A.K. SenGupta, *Ion Exchange and Solvent Extraction: A Series of Advances, volume 18*, CRC Press, Boca Raton, 2007. ISBN 9780849373978.

Y. Tanaka, *Ion Exchange Membranes: Fundamentals and Applications*, Elsevier Science, 2007. ISBN 9780444519825.

E.M. Thurman and M.S. Mills, *Solid-phase Extraction: Principles and Practice*, Chemical Analysis Series Vol. 147, J. Wiley & Sons, New York, 1998. ISBN 047161422X.

J. Weiss, *Handbook of Ion Chromatography*, (3rd Edn, 2 volumes), VCH-Wiley Publishing, 2004. ISBN 9783527287017.

A.A. Zagarondi, *Ion Exchange Materials: Properties and Applications*, Elsevier Science, 2007. ISBN 9780080445526.

## LABORATORY TECHNIQUE AND THEORETICAL DISCUSSION

*Chemical Principles in the Laboratory*, 9<sup>th</sup> Edn, Thompson Learning UK, 2008. ISBN 9780495112884.

G.S. Coyne, *The Laboratory Companion: A Practical Guide to Materials, Equipment and Technique*, Wiley-Interscience, New York, 1997. ISBN 0471184225.

J. Feist, R. Fessenden and J.S. Fessenden, *Organic Laboratory Techniques*, 3rd Edn, Brooks Cole Publishing Co, 2000. ISBN 9780534379810.

J.L. Hollenberg, J.M. Postma, Roberts J.L., *Chemistry in the Laboratory*, 6th Edn, W.H. Freeman & Co, NY, 2004. ISBN 9780716796060.

D.R. Lide, *CRC Handbook of Chemistry and Physics*. CRC Press, 88th Edn, Boca Raton, Florida, 2007. ISBN 9780849304880 and 9780849328787 (on cd-rom).

H.J.E. Loewenthal and E. Zass, *A Guide for the Perplexed Organic Experimentalist*, 2nd Edn, J. Wiley & Sons, Chichester, 1992. ISBN 0471935336.

W.L. Masterton, E.J. Slowinski, W. Wolsey, *Chemical Principles in the Laboratory*, Brooks-Cole Publishing Co, 2004. ISBN 9780534424534. [Slowinski, Wolsey and Masterton, *Eds—Chapters 1 to 43: Chemical Principles in the Laboratory*, Thompson Learning UK, 2006. ISBN (for Ch 43) 9780495371885.]

D.W. Mayo, S.S. Butcher, P.K. Trumper, R.M. Pike and S.S. Butcher, *Microscale Techniques for the Organic Laboratory*, J. Wiley & Sons, New York, 1991. ISBN 0471621927.

C.E. Meloan, *Chemical Separations, Principles, Techniques and Experiments*, J. Wiley & Sons, New York, 2000, ISBN 0471351970.

J. Mendham, R.C. Denney, J.D. Barnes and M.J.K. Thomas, *Vogel's Quantitative Chemical Analysis*, 6th Edn, Prentice Hall, Harlow, 2000. ISBN 0582226287.

J.R. Morig, C.N. Hammod and P.F. Schatz, *Techniques in Organic Chemistry: miniscale, standard taper microscale, and Williamson microscale*, W.H. Freeman & Co, NY, 2010. ISBN 9781429219563.

E. Prichard, *Practical Laboratory Skills Training Guide (complete set 5)*, Royal Society of Chemistry, 2003. ISBN 9780854044580.

P.F. Schatz, J.R. Mohrig and C.N. Hammond, *Techniques in Organic Chemistry*, 2nd Edn, Palgrave Macmillan, 2006. ISBN 9780716769354.

A.M. Schoffstall, B.A. Gaddis, M.L. Druelinger and M. Druelinger, *Organic Microscale and Miniscale Laboratory Experiments*, McGraw-Hill, Boston, 2000. ISBN 0072375493.

T. Toube, R. Keese and B. Brandle, *Practical Organic Synthesis: A Student's Guide*, J. Wiley & Sons Inc, 2006. ISBN 9780470029657.

K. Waldron and L. McGahey, *Chemistry of Everything: Laboratory Manual*, Prentice Hall Inc, 2006. ISBN 9780131875364.

S.A. Weiner, *Intro to Chemical Principles: A Laboratory Approach*, Thompson Learning UK, 2007. ISBN 9780495114796.

J.D. Wollins (Ed). *Inorganic Experiments*, VCH-Wiley Publishing, 1995. ISBN 9783527292530.

J.W. Zubrick, *The Organic Chemistry Survival Manual: A Student's Guide to Techniques*, 6th Edn, J. Wiley & Sons Inc, 2007. ISBN 9780470896051, also 9780470129326.

**MOLECULAR SIEVES**

J. Cejka, H. van Bekkum, A. Corma and F. Schueth, *Introduction to Zeolite Molecular Sieves*, Elsevier Science, 2007. ISBN 9780444530639.

H.G. Karge and J. Weitkamp (Eds), *Molecular Sieves, Characterization I*, Springer, 2004. ISBN 9783540643357.

J. Cejka, *Molecular Sieves: From Basic Research to Industrial Applications*, Elsevier Science, 2005. ISBN 0780444520821.

M. Lahmani, C. Bréchnignac and P. Houdy, *Nanomaterials and Nanochemistry*, Springer, 2007. ISBN 9783540729921.

J.A. Schwartz and C. Contescu (Eds), *Surfaces of Nanoparticles & Porous Materials*, Marcel Dekker Inc, 1999. ISBN 9780824719333.

R. Szostak, *Molecular Sieves: Principles of Synthesis and Identification*, Van Nostrand Reinhold, New York, 1989. ISBN 0442280238.

**REAGENTS AND PURITY**

J.I.G. Cadogan et al. (Eds), *Dictionary of Organic Compounds*, 6th Edn, Chapman and Hall Electronic Publ Div, 1996. ISBN (9 volume set) 0412540908.

S. Caron, *Practical Organic Chemistry: Proven Reactions and Synthetic Transformations*, J. Wiley & Sons Inc, 2008. ISBN 9780470037331.

W.J. Lennarz and M.D. Lane (Eds), *Encyclopedia of Biological Chemistry*, Elsevier Inc, 2004. ISBN (4 volumes set) 0124437109.

D.R. Lide (Editor-in-Chief), *CRC—Handbook of Chemistry and Physics*, 88th Edition, CRC Press Boca Raton, Florida 2007-2008, ISBN 0849304881.

M.J. O'Neil (Ed), *The Merck Index: An Encyclopedia of Chemicals, Drugs and Biologicals*, Merck & Co, NJ, 2006. ISBN 9780911971000.

L.A. Paquette (Ed-in-Chief), *Handbook of Reagents for Organic Synthesis*, J Wiley & Sons, Vol 1 1999, ISBN 47197924-4 — Vol 10 2008, ISBN 978047051811-3; cf <[www.wiley.com](http://www.wiley.com)>.

L.A. Paquette (Series Ed), *The Complete Collection of Organic Reagents On Line (up to 2006)*: <[www.interscience.wiley.com/eros/](http://www.interscience.wiley.com/eros/)>.

A.J. Pearson and W.R. Rouse, *Handbook of Reagents for Organic Synthesis*, 4 Vol. Set, J. Wiley & Sons, Chichester, 2001. ISBN 04719878911.

*Reagents Chemicals: American Chemical Society Specifications*, Official from January 1 2000 (Reagent Chemicals American Chemical Society Specification) 9th Edn, Washington DC (Sept. 1999). ISBN 0841236712.

J.G. Speight (Ed), *Lange's Handbook of Chemistry*, 16th Edn, McGraw-Hill Co Inc, 2005. ISBN 0071432205.

L.E. Smart, *Separation, Purification and Identification*, The Royal Society of Chemistry, London, 2002 (in paperback and CD-ROM). ISBN 0854046852.

G. Wypych (Ed.), *Solvents Database on CD Rom*, William Andrew Publishing, LLC 2001. ISBN 0815514638.

**SAFETY IN THE CHEMICAL LABORATORY**

M.-A. Armour, *Hazardous Laboratory Chemicals: Disposal Guide*, 2nd Edn, CRC Press, Boca Raton, 1996. ISBN 1566701082.

L. Bretherick, P.G. Urben and M.J. Pitt, *Bretherick's Handbook of Reactive Chemical Hazards*, 7th Edn, (2 volumes), Harcourt, 2006. ISBN 9780123739452.

N.P. Cheremisinoff, *Handbook of Hazardous Chemical Properties*, Butterworth-Heinemann, Oxford, 2000. ISBN 0750672099.

A.K. Furr (Ed), *CRC Handbook of Laboratory Safety* (5<sup>th</sup> Edn), CRC Press, Boca Raton, 2000. ISBN 0849325234.

W.Y. Garner and M.S. Barge, *Good Laboratory Practices: An Agrochemical Perspective*, American Chemical Society, Washington DC, 1988. ISBN 841214808.

S.K. Hall, *Chemical Safety in the Laboratory*, Lewis Publishers, 1993. ISBN 9780873718967.

G.J. Hathaway, *Chemical Hazards of the Workplace*, 4th Edn, J. Wiley & Sons, 1996. ISBN 0471287024.

*Journal of Chemical Health and Safety*. Elsevier, ISSN 1871-5532.

R.E. Lenga (Ed.), *The Sigma-Aldrich Library of Chemical Safety Data*, 3 Volume set, Aldrich Chemical Co., Milwaukee, 1988. ISBN 0941633160.

R.E. Lenga, K.L. Votoupal (Eds), *The Sigma-Aldrich Library of Regulatory and Safety Data*, 3 Volume set, Aldrich Chemical Co., Milwaukee, 1993. ISBN 0941633357.

R.J. Lewis Sr., *Sax's Dangerous Properties of Industrial Materials*, 10th Edn, 3 Vol. Set, J. Wiley & Sons, New York, 1999. ISBN 0471354074.

S.G. Luxon (Ed.), *Hazards in the Chemical Laboratory*, 5th Edn, Royal Society of Chemistry, 1992. ISBN 0851862292.

R.J. Lewis Sr., *Hazardous Chemicals Desk Reference*, 5th Edn, J. Wiley & Sons, 2002. ISBN 0471441651.

NIOSH – Pocket Guide to Chemical Hazards, (9)RS (ISBN 9781590425862), free download from the internet.

*Organic Process Research & Development — Annual Reviews*.

P. Patniak, *A Comprehensive Guide to the Hazardous Properties of Chemical Substances*, 2nd Edn, J. Wiley & Sons, 1999. ISBN 0471291757.

*Safety Sense: A Laboratory Guide*, Cold Spring Harbor Lab Press, 2007. ISBN 9780879697839.

Waste Disposal: "Prudent Practices in the Laboratory: Handling and Disposal of Chemicals" National Academy Press; Washington, DC, 1995, ISBN-10: 0309052297, ISBN-13: 978309052290. [This book can be read online, free.]

**SOLVENTS, SOLVENT EXTRACTION AND DISTRIBUTION**

D.R. Lide, *Handbook of Organic Solvents*, CRC Press, Florida, 1995. ISBN 0849389305.

R.P. Pohanish and S.A. Greene, *Rapid Guide to Chemical Incompatibilities*, J. Wiley & Sons, New York, 1997. ISBN 0471288020.

J. Rydberg, M. Cox and C. Musikas, *Solvent Extraction: Principles and Practice*, Marcel Dekker Inc, 2004. ISBN 9780824750633.

I.M. Smallwood, *Solvent Recovery Handbook*, Blackwell Science Publ Ltd, 2001. ISBN 9780632056477.

## SPECTROSCOPY

S. Becker, *Inorganic Mass Spectrometry: Principles and Applications*. J. Wiley & Sons Inc, 2007. ISBN 9780470012000.

S. Berger and S. Braun, *200 and More NMR Experiments*, VCH-Wiley Publishing, 2004. ISBN 9783527310678.

E. Breitmaier, *Structure Elucidation by NMR in Organic Chemistry: A Practical Guide*, J. Wiley & Sons, New York, 1993. ISBN 0471937452, 0471933813.

J.A.C. Broekaert, *Analytical Atomic Spectrometry with Flames and Plasmas*, VCH-Wiley Publishing, 2005. ISBN 9783527312825.

B. Henderson and G.F. Imbusch, *Optical Spectroscopy of Inorganic Solids*, Oxford University Press UK, 2006. ISBN 9780199298642.

W. Henderson and J.S. McIndoe, *Mass Spectrometry of Inorganic and Organometallic Compounds: Tools-Techniques-Tips*, J Wiley & Sons Inc, 2005. ISBN 9780470014318.

M. Hesse, H. Meier and B. Zeeh, *Spectroscopic Methods in Organic Chemistry* (translated by Dunmer & Murray) 2nd Edn, George Theme Verlag, 2007. ISBN 9783131060426 (Stuttgart) and 9781588904881 (New York).

De Hoffmann and V. Stroobant, *Mass Spectrometry: Principles and Applications*, J Wiley & Sons Inc, 2007. ISBN 9780470033111.

J.A. Iggo, *NMR Spectroscopy in Inorganic Chemistry*, Oxford University Press UK, 1999 (reprinted 2000). ISBN 9780198558903

K. Nakamoto, *Infrared and Raman Spectra of Inorganic and Coordination Compounds*, 2 Vol. Set, 5th Edn, J. Wiley & Sons, New York, 1997. ISBN 0471194069.

C.J. Pouchert, *The Aldrich Library of Infrared Spectra*, 3rd Edn, Aldrich Chemical Co., Milwaukee, 1981.

C.J. Pouchert and J. Behnke, *The Aldrich Library of  $^{13}\text{C}$  and  $^1\text{H}$  FT NMR Spectra*, 3 Volume set, Aldrich Chemical Co., Milwaukee, 1993. ISBN 0941633349.

J. Sole and L. Bausa, *An Introduction to the Optical Spectroscopy of Inorganic Solids*, J. Wiley & Sons Inc, 2005. ISBN 9780470868850.

T.J. Wenzel, *Discrimination of Chiral Compounds using NMR Spectroscopy*, Wiley-Interscience, New York 2007. ISBN 978047176352-9

D. Williams and I. Fleming, *Spectroscopy in Organic Chemistry*, McGraw-Hill Publishing Co UK, 2008. ISBN 9780077118129.

## THIN LAYER CHROMATOGRAPHY

B. Fried, J. Sherma, *Practical Thin-Layer Chromatography: A MultiDisciplinary Approach*, CRC Press, Boca Raton, 1996. ISBN 9780849326608.

B. Fried and J. Sherma, *Thin Layer Chromatography*, Marcel Dekker, New York, 1999. ISBN 0824702220.

E. Hahn-Dienstrop, *Applied Thin Layer Chromatography: Best Practice and Avoidance of Mistakes*, VCH Verlagsgesellschaft MbH, Weinheim, 2000. ISBN 3527298398.

H. Jork (Ed.), *Thin Layer Chromatography: Reagents and Detection Methods*, Vol. 1a, 1b, J. Wiley & Sons, 1989 (Vol. 1a) and 1994 (Vol. 1b). ISBN 3527278346 (Vol. 1a) and 352728205X (Vol. 1b).

T. Kowalska and J. Sherma, *Preparative Thin Layer Chromatography*, CRC Taylor & Francis Press, 2006. ISBN 10: 0-8493-4039-X, 13: 97804934390.

T. Kowalska and J. Sherma, *Thin Layer Chromatography in Chiral Separations and Analysis*, CRC Press, 2007. ISBN 978084943698.

# CHAPTER 2

## CHEMICAL METHODS USED IN PURIFICATION

### GENERAL REMARKS

Greater selectivity in purification can often be achieved by making use of differences in chemical properties between the substance to be purified and the contaminants. Unwanted metal ions may be removed by precipitation in the presence of a *collector* (see below). Sodium borohydride and other metal hydrides transform organic peroxides and carbonyl-containing impurities such as aldehydes and ketones in alcohols and ethers. Many classes of organic chemicals can be purified by conversion into suitable derivatives, followed by regeneration. This chapter describes relevant procedures.

### REMOVAL OF TRACES OF METALS FROM REAGENTS

#### METAL IMPURITIES

The presence of metal contaminants in reagents may sometimes affect the chemical or biochemical outcomes of an experiment. In these cases, it is necessary to purify the reagents used. Metal (and other) impurities can be determined qualitatively and quantitatively by atomic absorption spectroscopy (AAA), x-ray photoelectron spectroscopy (XPS), various mass spectrometric methods and/or inductively coupled plasma mass spectrometry (ICP-MS) (see Chapter 1, Question of Purity) and the required purification procedures can be formulated. Metal impurities in organic compounds are usually in the form of ionic salts or complexes with organic compounds and very rarely in the form of free metal. If they are present in the latter form then they can be removed by crystallising the organic compound (whereby the insoluble metal can be removed by filtration), or by distillation in which case the metal remains behind with the residue in the distilling flask. If the impurities are in the ionic or complex forms, then extraction of the organic compound in a suitable organic solvent with aqueous acidic or alkaline solutions will reduce their concentration to acceptable levels.

When the metal impurities are present in inorganic compounds as in metals or metal salts, then advantage of the differences in chemical properties should be taken. Properties of the impurities like the solubility, the solubility product (product of the metal ion and the counter-ion concentrations), the stability constants of the metal complexes with organic complexing agents and their solubilities in organic solvents should be considered. Alternatively the impurities can be masked by the addition of complexing agents which could lower the concentration of the metal ion impurities to such low levels that they would not interfere with the main compound (see **complexation** below). Specific procedures and examples are provided below.

#### DISTILLATION

Reagents such as water, ammonia, hydrochloric acid, nitric acid, perchloric acid, and sulfuric acid can be purified *via* distillation (preferably under reduced pressure and particularly with perchloric acid) using an all-glass still. Isothermal distillation is convenient for ammonia: a beaker containing concentrated ammonia is placed alongside a beaker of distilled water for several days in an empty desiccator so that some of the ammonia distils over into the water. The redistilled ammonia should be kept in polyethylene or paraffin-waxed bottles. Hydrochloric acid can be purified in the same way. To ensure the absence of metal contaminants from some salts (e.g. ammonium acetate), it may be more expedient to synthesise the salts using distilled components rather than to attempt to purify the salts themselves.

### SCAVENGER RESINS AND OTHER SUPPORTS

There is now an extensive range of supported reactants that use resins, silica, carbons etc, to clean up reactions prior to final purification which is gaining favour in the laboratory. [See section on "Scavenger Resins" in Chapter 3, at the end of the sections on "Preparation of other adsorbents", "FPLC" and "HPLC".]

### USE OF ION-EXCHANGE RESINS

Application of ion-exchange columns has greatly facilitated the removal of heavy metal ions such as  $\text{Cu}^{2+}$ ,  $\text{Zn}^{2+}$  and  $\text{Pb}^{2+}$  from aqueous solutions of many reagents. Thus, sodium salts and sodium hydroxide can be purified by passage through a column of a cation-exchange resin in its sodium form, prepared by washing the resin with 0.1M aqueous NaOH then washing with water until the pH of the effluent is  $\sim 7$ . Similarly, for acids, a resin in its  $\text{H}^+$  form [prepared by washing the column with 0.1M aqueous mineral acid ( $\text{HCl}$ ,  $\text{H}_2\text{SO}_4$ ) followed by thorough washing with water until the effluent has pH  $\sim 7$  is used]. In some cases, where metals form anionic complexes, they can be removed by passage through an anion-exchange resin. Iron in hydrochloric acid solution can be removed in this way.

Ion-exchange resins are also useful for demineralising biochemical preparations such as proteins. Removal of metal ions from protein solutions using polystyrene-based resins, however, may lead to protein denaturation. This difficulty may be avoided by using a weakly acidic cation exchanger such as Bio-Rex 70.

Heavy metal contamination of pH buffers can be removed by passage of the solutions through a Chelex X-100 column. For example when a solution of 0.02M HEPES [4-(2-Hydroxyethyl)piperazine-1-ethanesulfonic acid] containing 0.2M KCl (1L, pH 7.5) alone or with calmodulin, is passed through a column of Chelex X-100 ( $\sim 60\text{g}$ ) in the  $\text{K}^+$  form, the level of  $\text{Ca}^{2+}$  ions falls to less than  $2 \times 10^{-7}$  M as shown by atomic absorption spectroscopy. Such solutions should be stored in polyethylene containers that have been washed with boiling deionised water (5 minutes) and rinsed several times with deionised water. TES [*N,N,N',N'*-Tetraethylsulfamide] and TRIS [Tris-(hydroxymethyl)aminomethane] have been similarly decontaminated from metal ions.

Water, with very low concentrations of ionic impurities (and approaching conductivity standards), is very readily obtained by percolation through alternate columns of cation- and anion-exchange resins, or through a mixed-bed resin, and many commercial devices are available for this purpose. For some applications, this method is unsatisfactory because the final deionised water may contain traces of organic material after passage through the columns. However, organic matter can be removed by using yet another special column in series for this purpose (see Water Purification Systems-Pure and Ultrapure in <[www.millipore.com](http://www.millipore.com)>).

### PRECIPITATION

In removing traces of impurities by precipitation, it is necessary to include a material to act as a *collector* of the precipitated substance so as to facilitate its removal by filtration or decantation. The following are a few examples.

#### Removal of lead contaminants

Aqueous hydrofluoric acid can be freed from lead by adding 1ml of 10% strontium chloride per 100ml of acid, lead being co-precipitated as lead fluoride with the strontium fluoride. If the hydrofluoric acid is decanted from the precipitate and the process repeated, the final lead content in the acid is less than 0.003ppm. Similarly, lead can be precipitated from a nearly saturated sodium carbonate solution by adding 10% strontium chloride dropwise (1–2ml per 100ml) followed by filtration. (If the sodium carbonate is required as a solid, the solution can be evaporated to dryness in a platinum dish.) Removal of lead from potassium chloride uses precipitation as lead sulfide by bubbling  $\text{H}_2\text{S}$ , followed, after filtration, by evaporation and recrystallisation of the potassium chloride.

#### Removal of iron contaminants

Iron contaminants have been removed from potassium thiocyanate solutions by adding a slight excess of an aluminium salt, then precipitating aluminum and iron as their hydroxides by adding a few drops of ammonia. Iron is also carried down on the hydrated manganese dioxide precipitate formed in cadmium chloride or cadmium sulfate solutions by adding 0.5% aqueous potassium permanganate (0.5ml per 100ml of solution), sufficient ammonia to give a slight precipitate, and 1ml of ethanol. The solution is heated to boiling to coagulate the precipitate, then filtered. Ferrous ion can be removed from copper solutions by adding some hydrogen peroxide to the solution to oxidise the iron, followed by precipitation of ferric hydroxide by adding a small amount of sodium hydroxide.

### Removal of other metal contaminants

Traces of calcium can be removed from solutions of sodium salts by precipitation at pH 9.5–10 as the 8-hydroxyquinolate. The excess 8-hydroxyquinoline acts as a *collector* and is extracted out with an organic solvent.

### EXTRACTION

In some cases, a simple solvent extraction is sufficient to remove a particular impurity. For example, traces of gallium can be removed from titanous chloride in hydrochloric acid by extraction with diisopropyl ether. Similarly, ferric chloride can be removed from aluminium chloride solutions containing hydrochloric acid by extraction with diethyl ether. Usually, however, it is necessary to extract an undesired metal with an organic solvent in the presence of a suitable complexing agent such as dithizone (diphenylthiocarbazone) or sodium diethyl dithiocarbamate. When the former is used, weakly alkaline solutions of the substance containing the metal impurity are extracted with dithizone in chloroform (at about 25mg/L of chloroform) or carbon tetrachloride until the colour of some fresh dithizone solution remains unchanged after shaking. Dithizone complexes metals more strongly in weakly alkaline solutions. Excess dithizone in the aqueous medium is removed by extracting with the pure solvent (chloroform or carbon tetrachloride), the last traces of which, in turn, are removed by aeration. This method has been used to remove metal impurities from aqueous solutions of ammonium hydrogen citrate, potassium bromide, potassium cyanide, sodium acetate and sodium citrate. The advantage of dithizone for such a purpose lies in the wide range of metals with which it combines under these conditions. 8-Hydroxyquinoline (oxine) can also be used in this way. Sodium diethyl dithiocarbamate has been used to remove metals from aqueous hydroxylamine hydrochloride (made just alkaline to thymol blue by adding ammonia) from copper and other heavy metals by repeated extraction with chloroform until no more diethyl dithiocarbamate remained in the solution (which was then acidified to thymol blue by adding hydrochloric acid).

### COMPLEXATION

Although not strictly a removal of an impurity, addition of a suitable complexing agent such as ethylenediaminetetraacetic acid often overcomes the undesirable effects of contaminating metal ions by reducing the concentrations of the free metal species to very low levels, i.e. sequestering metal ions by complexation. For a detailed discussion of this *masking*, see *Masking and Demasking of Chemical Reactions*, D.D.Perrin, Wiley-Interscience, New York, 1970.

## USE OF METAL HYDRIDES

This group of reagents is commercially available in large quantities; some of its members—notably lithium aluminium hydride ( $\text{LiAlH}_4$ ), calcium hydride ( $\text{CaH}_2$ ), sodium borohydride ( $\text{NaBH}_4$ ) and potassium borohydride ( $\text{KBH}_4$ )-have found widespread use in the purification of chemicals.

### LITHIUM ALUMINIUM HYDRIDE

This solid is stable at room temperature and is soluble in ether-type solvents. It reacts violently with water, liberating hydrogen, and is a powerful drying and reducing agent for organic compounds. It reduces aldehydes, ketones, esters, carboxylic acids, peroxides, acid anhydrides and acid chlorides to the corresponding alcohols. Similarly, amides, nitriles, aldimines and aliphatic nitro compounds yield amines, while aromatic nitro compounds are converted to azo compounds. For this reason it finds extensive application in purifying organic chemical substances by the removal of water and carbonyl containing impurities as well as peroxides formed by autoxidation. Reactions can generally be carried out at room temperature, or in refluxing diethyl ether, at atmospheric pressure. *When drying organic liquids with this reagent it is important that the concentration of water in the liquid is below 0.1% - otherwise a violent reaction or EXPLOSION may occur.  $\text{LiAlH}_4$  should be added cautiously to a cooled solution of organic liquid in a flask equipped with a reflux condenser.*

### CALCIUM HYDRIDE

This powerful drying agent is suitable for use with hydrogen, argon, helium, nitrogen, hydrocarbons, chlorinated hydrocarbons, esters and higher alcohols.

### SODIUM BOROHYDRIDE

This solid, which is stable in dry air up to 300°, is a less powerful reducing agent than lithium aluminium hydride, from which it differs also by being soluble in hydroxylic solvents and to a lesser extent in ether-type solvents. Sodium borohydride forms a dihydrate melting at 36–37°, and its aqueous solutions decompose slowly unless stabilised to above pH 9 by alkali. (For example, a useful sodium borohydride solution is one that is nearly saturated at 30–40° and containing 0.2% sodium hydroxide.) Its solubility in water is 25, 55 and 88g per 100ml of water at 0, 25 and 60°, respectively. Boiling or acidification rapidly decomposes aqueous sodium borohydride solutions. The reagent, available either as a hygroscopic solid or as an aqueous sodium hydroxide solution, is useful as a water soluble reducing agent for aldehydes, ketones and organic peroxides. This explains its use for the removal of carbonyl-containing impurities and peroxides from alcohols, polyols, esters, polyesters, amino alcohols, olefins, chlorinated hydrocarbons, ethers, polyethers, amines (including aniline), polyamines and aliphatic sulfonates.

Purifications using sodium borohydride can be carried out conveniently using alkaline aqueous or methanolic solutions of sodium borohydride, allowing the reaction mixture to stand at room temperature for several hours. Other solvents that can be used with this reagent include isopropyl alcohol (without alkali), amines (including liquid ammonia, in which its solubility is 104g per 100g of ammonia at 25°, and ethylene diamine), diglyme, formamide, dimethylformamide and tetrahydrofurfuryl alcohol. Alternatively, the material to be purified can be percolated through a column of the borohydride. In the absence of water, sodium borohydride solutions in organic solvents such as dioxane or amines decompose only very slowly at room temperature. Treatment of ethers with sodium borohydride appears to inhibit peroxide formation.

### POTASSIUM BOROHYDRIDE

Potassium borohydride is similar in properties and reactions to sodium borohydride, and can similarly be used as a reducing agent for removing aldehydes, ketones and organic peroxides. It is non-hygroscopic and can be used in water, ethanol, methanol or water-alcohol mixtures, provided some alkali is added to minimise decomposition, but it is somewhat less soluble than sodium borohydride in most solvents. For example, the solubility of potassium borohydride in water at 25° is 19g per 100ml of water (as compared to 55g of sodium borohydride).

## PURIFICATION *via* DERIVATIVES

Relatively few derivatives of organic substances are suitable for use as aids to purification. This is because of the difficulty in regenerating the starting material. For this reason, we list below the common methods of preparation of derivatives that can be used in this way.

Whether or not any of these derivatives is likely to be satisfactory for the use of any particular case will depend on the degree of difference in properties, such as solubility, volatility or melting point, between the starting material, its derivative and likely impurities, as well as on the ease with which the substance can be recovered. Purification *via* a derivative is likely to be of most use when the quantity of pure material that is required is not too large. Where large quantities (for example, more than 50g) are available, it is usually more economical to purify the material directly (for example, in distillations and recrystallisations).

**The most generally useful purifications *via* derivatives are as follows.**

### ALCOHOLS

Aliphatic or aromatic alcohols are converted to solid esters. *p*-Nitrobenzoates are examples of convenient esters to form because of their sharp melting points, and the ease with which they can be recrystallised as well as the ease with which the parent alcohol can be recovered. The *p*-nitrobenzoyl chloride used in the esterification is prepared by refluxing dry *p*-nitrobenzoic acid with a 3 molar excess of thionyl chloride for 30 minutes on a steam bath (*in a fume cupboard*). The solution is cooled slightly and the excess thionyl chloride is distilled off under vacuum, keeping the temperature below 40°. Dry toluene is added to the residue in the flask, then distilled off under vacuum, the process being repeated two or three times to ensure complete removal of thionyl chloride, hydrogen chloride and sulfur dioxide. (This freshly prepared *p*-nitrobenzoyl chloride cannot be stored without decomposition; it should be used directly.) A solution of the acid chloride (1mol) in dry toluene or alcohol-free chloroform (distilled from P<sub>2</sub>O<sub>5</sub> or by passage through an activated Al<sub>2</sub>O<sub>3</sub> column) under a reflux condense is cooled in an ice bath while the alcohol (1mol), with or without a solvent (preferably miscible

with toluene or alcohol-free chloroform), is added dropwise to it. When addition is over and the reaction subsides, the mixture is refluxed for 30 minutes and the solvent is removed under reduced pressure. The solid ester is then recrystallised to constant melting point from toluene, acetone, low boiling point petroleum ether or mixtures of these, but not from alcohols (due to unwanted trans-esterification).

Hydrolysis of the ester is achieved by refluxing in aqueous N or 2N NaOH solution until the insoluble ester dissolves. The solution is then cooled, and the alcohol is extracted into a suitable solvent, e.g. ether, toluene or alcohol-free chloroform. The extract is dried ( $\text{CaSO}_4$ ,  $\text{MgSO}_4$ ) and distilled, then fractionally distilled if liquid or recrystallised if solid. (The *p*-nitrobenzoic acid can be recovered by acidification of the aqueous layer.) In most cases where the alcohol to be purified can be readily extracted from ethanol, the hydrolysis of the ester is best achieved with N or 2N ethanolic NaOH or 85% aqueous ethanolic N NaOH. The former is prepared by dissolving the necessary alkali in a minimum volume of water and diluting with absolute alcohol. The ethanolic solution is refluxed for one to two hours and hydrolysis is complete when an aliquot gives a clear solution on dilution with four or five times its volume of water. The bulk of the ethanol is distilled off and the residue is extracted as above. Alternatively, use can be made of ester formation with benzoic acid, toluic acid or 3,5-dinitrobenzoic acid, by the above method.

Other derivatives can be prepared by reaction of the alcohol with an acid anhydride. For example, phthalic or 3-nitrophthalic anhydride (1 mol) and the alcohol (1 mol) are refluxed for half to one hour in a non-hydroxylic solvent, e.g. toluene or alcohol-free chloroform, and then cooled. The phthalate ester crystallises out, is precipitated by the addition of low boiling petroleum ether or is isolated by evaporation of the solvent. It is recrystallised from water, 50% aqueous ethanol, toluene or low boiling petroleum ether. Such an ester has a characteristic melting point and the alcohol can be recovered by acid or alkaline hydrolysis.

### ALDEHYDES

The best derivative from which an aldehyde can be recovered readily is its bisulfite addition compound, the main disadvantage being the lack of a sharp melting point. The aldehyde (sometimes in ethanol) is shaken with a cold saturated solution of sodium bisulfite until no more solid adduct separates. The adduct is filtered off, washed with a little water, followed by alcohol. A better reagent to use is a freshly prepared saturated aqueous sodium bisulfite solution to which 75% ethanol is added to near-saturation. (Water may have to be added dropwise to render this solution clear.) With this reagent the aldehyde need not be dissolved separately in alcohol and the adduct is finally washed with alcohol. The aldehyde is recovered by dissolving the adduct in the least volume of water and adding an equivalent quantity of sodium carbonate (not sodium hydroxide) or concentrated hydrochloric acid to react with the bisulfite, followed by steam distillation or solvent extraction.

Other derivatives that can be prepared are the Schiff bases and semicarbazones. Condensation of the aldehyde with an equivalent of primary aromatic amine yields the Schiff base, for example aniline at 100° for 10–30 minutes.

Semicarbazones are prepared by dissolving semicarbazide hydrochloride (*ca* 1g) and sodium acetate (*ca* 1.5g) in water (8–10ml) and adding the aldehyde or ketone (0.5–1g) with stirring. The semicarbazone crystallises out and is recrystallised from ethanol or aqueous ethanol. These are hydrolysed by steam distillation in the presence of oxalic acid or better by exchange with pyruvic acid (Hershberg *J Org Chem* **13** 542 1948) [see entry under Ketones].

### AMINES

#### Picrates

The most versatile derivative from which the free base can be readily recovered is the picrate. This is very satisfactory for primary and secondary aliphatic amines and aromatic amines and is particularly so for heterocyclic bases. The amine, dissolved in water or alcohol, is treated with excess of a saturated solution of picric acid in water or alcohol, respectively, until separation of the picrate is complete. If separation does not occur, the solution is stirred vigorously and warmed for a few minutes, or diluted with a solvent in which the picrate is insoluble. Thus, a solution of the amine and picric acid in ethanol can be treated with petroleum ether to precipitate the picrate. Alternatively, the amine can be dissolved in alcohol and aqueous picric acid added. The picrate is filtered off, washed with water or ethanol and recrystallised from boiling water, ethanol, methanol, aqueous ethanol, methanol or chloroform. The solubility of picric acid in water and ethanol is 1.4 and 6.23% respectively at 20°.

It is not advisable to store large quantities of picrates for long periods, *particularly when they are dry due to their potential EXPLOSIVE nature*. Also this method is not advised when large quantities of amines (e.g. >25g) are to be purified. The free base should be recovered as soon as possible. The picrate is suspended in an

excess of 2N aqueous NaOH and warmed a little. Because of the limited solubility of sodium picrate, excess hot water must be added. Alternatively, because of the greater solubility of lithium picrate, aqueous 10% lithium hydroxide solution can be used. The solution is cooled, the amine is extracted with a suitable solvent such as diethyl ether or toluene, washed with 5N NaOH until the alkaline solution remains colourless, then with water, and the extract is dried with anhydrous sodium carbonate. The solvent is distilled off and the amine is fractionally distilled (under reduced pressure if necessary) or recrystallised.

If the amines are required as their hydrochlorides, picrates can often be decomposed by suspending them in acetone and adding two equivalents of 10N HCl. The hydrochloride of the base is filtered off, leaving the picric acid in the acetone. Dowex No 1 anion-exchange resin in the chloride form is useful for changing solutions of the more soluble picrates (for example, of adenosine) into solutions of their hydrochlorides, from which sodium hydroxide precipitates the free base.

#### Salts

Amines can also be purified *via* their salts, e.g. hydrochlorides. A solution of the amine in dry toluene, diethyl ether, dichloromethane or chloroform is saturated with dry hydrogen chloride (generated by addition of concentrated sulfuric acid to dry sodium chloride, or to concentrated HCl followed by drying the gas through sulfuric acid, or the HCl gas is obtained from a hydrogen chloride cylinder) and the insoluble hydrochloride is filtered off and dissolved in water. The solution is made alkaline and the amine is extracted, as above. Hydrochlorides can also be prepared by dissolving the amine in ethanolic HCl and adding diethyl ether. Where hydrochlorides are too hygroscopic or too soluble for satisfactory isolation, other salts, e.g. nitrate, sulfate, bisulfate or oxalate, can be used.

#### Double salts

The amine (1mol) is added to a solution of anhydrous zinc chloride (1mol) in concentrated HCl (42ml) in ethanol (200ml, or less depending on the solubility of the double salt). The solution is stirred for 1 hour and the precipitated salt is filtered off and recrystallised from ethanol. The free base is recovered by adding excess of 5-10N NaOH (to dissolve the zinc hydroxide that separates) and is steam distilled. Mercuric chloride (highly poisonous) in hot water can be used instead of zinc chloride and the salt is crystallised from 1% hydrochloric acid. Other double salts have been used, e.g. cuprous salts, but are not as convenient as the above salts.

#### N-Acetyl derivatives

Purification as their *N*-acetyl derivatives is satisfactory for primary, and to a limited extent, secondary amines. Tertiary amines are not acetylated. The base is refluxed with slightly more than one equivalent of acetic anhydride for half to one hour, cooled and poured into ice-cold water. The insoluble derivative is filtered off, dried, and recrystallised from water, ethanol, aqueous ethanol or benzene (CAUTION **toxic!**). The derivative can be hydrolysed to the parent amine by refluxing with 70% sulfuric acid for a half to one hour. The solution is cooled, poured onto ice, and made alkaline. The amine is steam distilled or extracted as above. Alkaline hydrolysis is very slow.

#### N-Tosyl derivatives

Primary and secondary amines are converted into their tosyl derivatives by mixing equimolar amounts of amine and *p*-toluenesulfonyl chloride in dry pyridine (*ca* 5-10mols) and allowing to stand at room temperature overnight. The solution is poured into ice-water and the pH adjusted to 2 with HCl. The solid derivative is filtered off, washed with water, dried (vacuum desiccator) and recrystallised from an alcohol or aqueous alcohol solution to a sharp melting point. The derivative is decomposed by dissolving in liquid ammonia (*fume cupboard*) and adding sodium metal (in small pieces with stirring) until the blue colour persists for 10-15 minutes. Ammonia is allowed to evaporate (*fume cupboard*), the residue treated with water and the solution checked that the pH is above 10. If the pH is below 10, then the solution has to be basified with 2N NaOH. The mixture is extracted with diethyl ether or toluene, the extract is dried ( $K_2CO_3$ ), evaporated and the residual amine recrystallised if solid or distilled if liquid.

## AROMATIC HYDROCARBONS

### Adducts

Aromatic hydrocarbons can be purified as their picrates using the procedures described for amines. Instead of picric acid, 1,3,5-trinitrobenzene or 2,4,7-trinitrofluorenone can also be used. In all these cases, following recrystallisation, the hydrocarbon can be isolated either as described for amines or by passing a solution of the adduct through an activated alumina column and eluting with toluene or petroleum ether. The picric acid and nitro compounds are more strongly adsorbed on the column.

### Sulfonation

Naphthalene, xylenes and alkyl benzenes can be purified by sulfonation with concentrated sulfuric acid and crystallisation of the sodium sulfonates. The hydrocarbon is distilled out of the mixture with superheated steam.

## CARBOXYLIC ACIDS

### 4-Bromophenacyl esters

A solution of the sodium salt of the acid is prepared. If the salt is not available, the acid is dissolved in an equivalent of aqueous NaOH and the pH adjusted to 8–9 with this base. A solution of one equivalent of 4-bromophenacyl bromide (for a monobasic acid, two equivalents for a dibasic acid, etc) in ten times its volume of ethanol is then added. The mixture is heated to boiling, and, if necessary, enough ethanol is added to clarify the solution which is then refluxed for half an hour to three hours depending on the number of carboxylic groups that have to be esterified. (One hour is generally sufficient for monocarboxylic acids.) On cooling, the ester should crystallise out. If it does not, then the solution is heated to boiling, and enough water is added to produce a slight turbidity. The solution is again cooled. The ester is collected, and recrystallised or fractionally distilled.

The ester is hydrolysed by refluxing for 1–2 hours with 1–5% of barium carbonate suspended in water or with aqueous sodium carbonate solution. The solution is cooled and extracted with diethyl ether, toluene or chloroform. It is then acidified and the acid is collected by filtration or extraction, and recrystallised or fractionally distilled. 4-Bromophenylphenacyl esters are used similarly.

*p*-Nitrobenzyl esters can be prepared in an analogous manner using the sodium salt of the acid and *p*-nitrobenzyl bromide. They are readily hydrolysed.

### Alkyl esters

Of the alkyl esters, methyl esters are the most useful because of their rapid hydrolysis. The acid is refluxed with one or two equivalents of methanol in excess alcohol-free chloroform (or dichloromethane) containing about 0.1g of *p*-toluenesulfonic acid (as catalyst), using a Dean-Stark apparatus. (The water formed by the esterification is carried away into the trap.) When the theoretical amount of water is collected in the trap, esterification is complete. The chloroform solution in the flask is washed with 5% aqueous sodium carbonate solution, then water, and dried over anhydrous sodium sulfate or magnesium sulfate. The chloroform is distilled off and the ester is fractionally distilled through an efficient column, or recrystallised if it is a solid. The ester is hydrolysed by refluxing with 5–10% aqueous NaOH solution until the insoluble ester has completely dissolved. The aqueous solution is concentrated a little by distillation to remove almost all of the methanol. It is then cooled and acidified. The acid is either extracted with diethyl ether, toluene or chloroform, or filtered off and isolated as above. Other methods for preparing esters are available, e.g. addition of an ethereal solution of diazomethane (yellow in colour, poisonous, use a fume cupboard. CARE: Use diazomethane with extreme care as the reagent is POISONOUS and HIGHLY explosive; special precautions MUST be used; see *Fieser and Fieser's Reagents for Organic Synthesis* 1 pp. 191–195 1967) to the acid which dissolves as the acid is esterified, liberating N<sub>2</sub> and the yellow colour. The methyl ester so produced is obtained by evaporating the ethereal solution.

### Salts

The most useful salt derivatives for carboxylic acids are the isothiuronium salts. These are prepared by mixing almost saturated solutions containing the acid (carefully neutralised with N NaOH using phenolphthalein indicator) then adding two drops of N HCl and an equimolar amount of *S*-benzylisothiuronium chloride in ethanol and filtering off the salt that crystallises out. After recrystallisation from water, alcohol or aqueous alcohol the salt is decomposed by suspending or dissolving in 2N HCl and extracting the carboxylic acid from aqueous solution into diethyl ether, chloroform or toluene.

## HYDROPEROXIDES

These can be converted to their sodium salts by precipitation below 30° with aqueous 25% NaOH. The salt is then decomposed by addition of solid (powdered) carbon dioxide and extracted with low-boiling petroleum ether. The solvent should be removed under reduced pressure below 20°. **The manipulation should be adequately shielded at all times to guard against EXPLOSIONS for the safety of the operator.**

## KETONES

### Bisulfite adduct

The adduct can be prepared and decomposed as described for aldehydes. Alternatively, because no Cannizzaro reaction is possible, it can also be decomposed with 0.5N NaOH.

### Semicarbazones

A powdered mixture of semicarbazide hydrochloride (1mol) and anhydrous sodium acetate (1.3mol) is dissolved in water by gentle warming. A solution of the ketone (1mol) in the minimum volume of ethanol needed to dissolve it is then added. The mixture is warmed on a water bath until separation of the semicarbazone is complete. The solution is cooled, and the solid filtered off. After washing with a little ethanol followed by water, it is recrystallised from ethanol or dilute aqueous ethanol. The derivative should have a characteristic melting point. The semicarbazone is decomposed by refluxing with excess of oxalic acid or with

aqueous sodium carbonate solution. The ketone (which steam distils) is distilled off. It is extracted or separated from the distillate (after saturating with NaCl), dried with CaSO<sub>4</sub> or MgSO<sub>4</sub> and fractionally distilled using an efficient column (under vacuum if necessary). [See entry under Aldehydes.]

### PHENOLS

The most satisfactory derivatives for phenols that are of low molecular weight or monohydric are the benzoate esters. (Their acetate esters are generally liquids or low-melting solids.) Acetates are more useful for high molecular weight and polyhydric phenols.

#### Benzoates

The phenol (1mol) in 5% aqueous NaOH is treated (while cooling) with benzoyl chloride (1mol) and the mixture is stirred in an ice bath until separation of the solid benzoyl derivative is complete. The derivative is filtered off, washed with alkali, then water, and dried (in a vacuum desiccator over NaOH). It is recrystallised from ethanol or dilute aqueous ethanol. The benzoylation can also be carried out in dry pyridine at low temperature (*ca* 0°) instead of in NaOH solution, finally pouring the mixture into water and collecting the solid as above. The ester is hydrolysed by refluxing in an alcohol (for example, ethanol, *n*-butanol) containing two or three equivalents of the alkoxide of the corresponding alcohol (for example, sodium ethoxide or sodium *n*-butoxide) and a few (*ca* 5–10) millilitres of water, for half an hour to three hours. When hydrolysis is complete, an aliquot will remain clear on dilution with four to five times its volume of water. Most of the solvent is distilled off. The residue is diluted with cold water and acidified, and the phenol is steam distilled. The latter is collected from the distillate, dried and either fractionally distilled or recrystallised. It can also be isolated by extraction from a slightly acidified (pH ~3) aqueous solution with diethyl ether.

#### Acetates

These can be prepared as for the benzoates using either acetic anhydride with 3N NaOH or acetyl chloride in pyridine. They are hydrolysed as described for the benzoates. This hydrolysis can also be carried out with aqueous 10% NaOH solution, completion of hydrolysis being indicated by the complete dissolution of the acetate in the aqueous alkaline solution. On steam distillation, acetic acid also distils off, but in these cases the phenols (see above) are invariably solids which can be filtered off and recrystallised.

### PHOSPHATE AND PHOSPHONATE ESTERS

These can be converted to their uranyl nitrate addition compounds. The crude or partially purified ester is saturated with uranyl nitrate solution and the adduct is filtered off. It is recrystallised from *n*-hexane, toluene or ethanol. For the more soluble members crystallisation from hexane using low temperatures (–40°) has been successful. The adduct is decomposed by shaking with sodium carbonate solution and water, the solvent is steam distilled (if hexane or toluene is used) and the ester is collected by filtration. Alternatively, after decomposition, the organic layer is separated, dried with CaCl<sub>2</sub> or BaO, filtered, and fractionally distilled under high vacuum.

### MISCELLANEOUS

Impurities can sometimes be removed by conversion to derivatives under conditions where the major component does not react or reacts much more slowly. For example, normal (straight-chain) paraffins can be freed from unsaturated and branched-chain components by taking advantage of the greater reactivity of the latter with chlorosulfonic acid or bromine. Similarly, the preferential nitration of aromatic hydrocarbons can be used to remove e.g. benzene or toluene from cyclohexane by shaking for several hours with a mixture of concentrated nitric acid (25%), sulfuric acid (58%), and water (17%).

## GENERAL METHODS FOR THE PURIFICATION OF CLASSES OF COMPOUNDS

Chapters 4, 5 and 6 list a large number of individual compounds, with a brief statement of how each one may be purified. For substances that are not included in these chapters the following procedures may prove helpful.

### PROCEDURES

If the laboratory worker does not know of a reference to the preparation of a commercially available substance, he/she may be able to make a reasonable guess at the synthetic methods used from published laboratory syntheses.

This information, in turn, can simplify the necessary purification steps by suggesting probable contaminants. Physical methods of purification depend largely on the melting and boiling points of the materials. For gases and low-boiling liquids use is commonly made of the *freeze-pump-thaw* procedure. Gas chromatography is also useful, especially for low-boiling point liquids. Liquids are usually purified by refluxing with drying agents, acids or bases, reducing agents, charcoal, etc., followed by fractional distillation under reduced pressure. For solids, general methods include fractional freezing of the melted material, taking the middle fraction. Another procedure is sublimation of the solid under reduced pressure. The other commonly used method for purifying solids is by recrystallisation from a solution in a suitable solvent, by cooling with or without the prior addition of a solvent in which the solute is not very soluble (see Chapter 1).

The nature of the procedure will depend to a large extent on the quantity of purified material that is required. For example, for small quantities (50–250mg) of a pure volatile liquid, preparative gas chromatography is probably the best method. Two passes through a suitable column may well be sufficient. Similarly, for smaller amounts (100–500 mg) of an organic solid, column chromatography is likely to be satisfactory, the eluate being collected as a number of separate fractions (*ca* 5–10 ml) which are examined by FT-IR, NMR or UV spectroscopy, TLC or by some other appropriate analytical technique. (For information on suitable adsorbents and eluents the texts referred to in the bibliography at the end of Chapters 1 and 2 should be consulted.) Preparative thin layer chromatography or HPLC, FC and HPFC can also be used successfully for purifying up to 500 mg of solid. The latter chromatographic techniques (see Chapter 1) are more and more commonly used procedures for the purification of small molecules as well as large molecules such as polypeptides and DNA.

Where larger quantities (upwards of 1g) are required, most of the impurities should be removed by preliminary treatments, such as solvent extraction, liquid-liquid partition, or conversion to a derivative (*vide supra*) which can be purified by crystallisation or fractional distillation before being reconverted to the starting material. The substance is then crystallised or distilled. If the final amounts must be in excess of 25 g, preparation of a derivative is sometimes omitted because of the cost involved. In all of the above cases, purification is likely to be more laborious if the impurity is an isomer or a derivative with closely similar physical properties.

### CRITERIA OF PURITY

Purification becomes meaningful only insofar as adequate tests of purity are applied: the higher the degree of purity that is sought, the more stringent these tests must be. For this, the experimenter has to resort, in the first place, to preliminary physical methods such as melting and boiling points, chromatographic and spectroscopic procedures which are described in detail in Chapter 1. If the material is an organic solid, its melting point should first be taken and compared with the recorded value. *Note that the melting points of most salts, organic or inorganic, are generally decomposition points and are not reliable criteria of purity.* As part of the preliminary examination, the sample might be examined by thin layer chromatography in several different solvent systems and in high enough concentrations to facilitate the detection of minor components. On the other hand, if the substance is a liquid, its boiling point should be measured. If, further, the boiling point of the liquid is too high, or it decomposes on heating, then its purity should be assessed by high pressure liquid chromatography. Liquids, especially volatile ones, can be studied very satisfactorily by gas chromatography, preferably using at least two different stationary and/or mobile phases. Spectroscopic methods, if facilities are available, such as atomic absorption spectroscopy (AAA), and inductively coupled plasma mass spectrometry (ICP-MS) are useful and sensitive methods for detecting metal impurities and the concentrations of metals and metal salts or complexes.

Application of these tests at successive steps will give a good indication of whether or not the purification is satisfactory and will also show when adequate purification has been achieved. Finally elemental analyses, e.g. of carbon, hydrogen, nitrogen, sulfur, metals etc., are very sensitive to impurities (other than with isomers), and are good criteria of purity.

There are certain requirements for purity of new compounds in most journals. This is especially so for samples which are shown to have biological activity. See instructions to authors for ACS journals especially in *J. Med. Chem* (see “Guidelines for Authors under Purity Criteria of Tested Compounds” <<http://pubs.acs.org/journal/jmcmr>>.)

## GENERAL PROCEDURES FOR THE PURIFICATION OF SOME CLASSES OF ORGANIC COMPOUNDS

In the general methods of purification described below, it is assumed that the impurities belong essentially to a class of compounds different from the one being purified. They are suggested for use in cases where substances are not listed in Chapters 4, 5 and the low-molecular-weight compounds in Chapter 6. In such cases, the experimenter is advised to employ them in conjunction with information given in these chapters for the purification of suitable analogues. Also, for a wider range of drying agents and the use of cartridges (e.g.  $\text{Na}_2\text{SO}_4$  for removal of  $\text{H}_2\text{O}$ , or Celite for removal of tar), solvents for extraction and solvents for recrystallisation, the reader is referred to Chapter 1. A common method of purification of organic compounds is to convert them to a suitable derivative which is purified (with the assumption that the impurity does not form a similar derivative, or if it does its properties are different), and then regenerate the original compound and purify it further. Various derivatives are described for different classes of compounds below, but many more can be considered, and the reader is referred to texts on "Protecting Groups" which describe ways of selectively protecting functional groups and facile means of deprotecting them, i.e. regenerating the unprotected group [P.J. Kocienski *Protecting Groups* Thieme International Publisher, 2005, ISBN 9783131356031; J.R. Hanson *Protecting Groups in Organic Synthesis* Sheffield Academic Press, 1999, ISBN 9781850759577, or J. Wiley & Sons Inc, 1999, ISBN 980632045068]. See Chapter 6 for general purification procedures used for macromolecules.

### ACETALS

These are generally diethyl or dimethyl acetal derivatives of aldehydes. They are more stable to alkali than to acids. Their common impurities are the corresponding alcohol, aldehyde and water. Drying with sodium wire removes alcohols and water, and polymerizes aldehydes so that, after decantation, the acetal can be fractionally distilled. In cases where the use of sodium is too drastic, aldehydes can be removed by shaking with alkaline hydrogen peroxide solution and the acetal is dried with sodium carbonate or potassium carbonate. Residual water and alcohols (up to *n*-propyl) can be removed with Linde type 4A molecular sieves. The acetal is then filtered and fractionally distilled. Solid acetals (i.e. acetals of high-molecular-weight aldehydes) are generally low-melting and can be recrystallised from low-boiling petroleum ether, toluene or a mixture of both.

### ACIDS

#### Carboxylic acids

Liquid carboxylic acids are first freed from neutral and basic impurities by dissolving them in aqueous alkali and extracting with diethyl ether. (The pH of the solution should be at least three units above the  $\text{pK}_a$  of the acid, see  $\text{pK}$  in Chapter 1). The aqueous phase is then acidified to a pH at least three units below the  $\text{pK}_a$  of the acid and again extracted with ether. It is quite unnecessary to add large excesses of mineral acid (e.g. HCl) to liberate the organic acid, as mineral acids dissolve appreciably in organic solvents such as diethyl ether. The extract is dried with magnesium sulfate or sodium sulfate and the ether is distilled off. The acid is fractionally distilled through an efficient column. It can be further purified by conversion to its methyl or ethyl ester (*vide supra*) which is then fractionally distilled. Hydrolysis yields the original acid which is again purified as above.

Acids that are solids can be purified in this way, except that distillation is replaced by repeated crystallisation (preferable from at least two different solvents such as water, alcohol or aqueous alcohol, toluene, toluene/petroleum ether or acetic acid.) Water-insoluble acids can be partially purified by dissolution in N sodium hydroxide solution and precipitation with dilute mineral acid. If the acid is required to be free from sodium ions, then it is better to dissolve the acid in hot N ammonia, heat to *ca* 80°, adding slightly more than an equal volume of N formic acid and allowing to cool slowly for crystallisation. Any ammonia, formic acid or ammonium formate that adhere to the acid are removed when the acid is dried in a vacuum — these are volatile. Cartridges and columns are available (e.g. the Isolute SCX-2 made of polysulfonic acid bonded to silica as an ion exchange column developed by Biotage Inc (<[www.biotage.com](http://www.biotage.com)>), particularly for the purification of acids using flash chromatography.

The separation and purification of naturally occurring fatty acids, based on distillation, salt solubility and low temperature crystallisation, are described by K.S. Markley (Ed.), *Fatty Acids*, 2nd Edn, part 3, Chap. 20, Interscience, New York, 1964, see also N. Reavley *Essential Fatty Acids* Book Media Publ, 2002, ISBN 9780958157643; G. Grati and K. Sato (Eds) *Crystallisation and Polymorphism of Fats and Fatty Acids* Marcel Dekker, 1988, ISBN 9780824778750.

Aromatic carboxylic acids can be purified by conversion to their sodium salts, recrystallisation from hot water, and reconversion to the free acids.

### Sulfonic acids

The low solubility of sulfonic acids in organic solvents and their high solubility in water makes necessary a treatment different from that for carboxylic acids. Sulfonic acids are strong acids, they have the tendency to hydrate, and many of them contain water of crystallisation. The lower-melting and liquid acids can generally be purified with only slight decomposition by fractional distillation, preferably under reduced pressure. A common impurity is sulfuric acid, but this can be removed by recrystallisation from concentrated aqueous solutions. The wet acid can be dried by azeotropic removal of water with toluene, followed by distillation. The higher-melting acids, or acids that melt with decomposition, can be recrystallised from water or, occasionally, from ethanol. For a typical purification of aromatic sulfonic acids using their barium salts refer to benzenesulfonic acid in the "Aromatic Compounds" section in Chapter 4.

### Sulfinic acids

These acids are less stable, less soluble and less acidic than the corresponding sulfonic acids. The common impurities are the respective sulfonyl chlorides from which they have been prepared, and the thioisulfonates (neutral) and sulfonic acids into which they decompose. The first two of these can be removed by solvent extraction from an alkaline solution of the acid. On acidification of an alkaline solution, the sulfinic acid crystallises out leaving the sulfonic acid behind. The lower molecular weight members are isolated as their metal (e.g. ferric) salts, but the higher members can be crystallised from water (made slightly acidic), or alcohol.

### ACID CHLORIDES

The corresponding acid and hydrogen chloride are the most likely impurities. Usually these can be removed by efficient fractional distillation. Where acid chlorides are not readily hydrolysed (e.g. aryl sulfonyl chlorides) the compound can be freed from contaminants by dissolving in a suitable solvent such as alcohol-free chloroform, dry toluene or petroleum ether and shaking with dilute sodium bicarbonate solution. The organic phase is then washed with water, dried with anhydrous sodium sulfate or magnesium sulfate, and distilled or recrystallised. This procedure is *hazardous* with readily hydrolysable *carboxylic* acid chlorides such as acetyl chloride and benzoyl chloride. Solid acid chlorides should be thoroughly dried *in vacuo* over strong drying agents and are satisfactorily recrystallised from toluene, toluene-petroleum ether, petroleum ethers, alcohol-free chloroform/toluene, and, occasionally, from dry diethyl ether. Hydroxylic or basic solvents should be strictly avoided. *All operations should be carried out in a fume cupboard because of the irritant nature of these compounds which also attack the skin.*

### ALCOHOLS

#### Monohydric alcohols

The common impurities in alcohols are aldehydes or ketones, and water. [*Ethanol* in Chapter 4 is typical.] Aldehydes and ketones can be removed by adding a small amount of sodium metal and refluxing for 2 hours, followed by distillation. Water can be removed in a similar way but it is preferable to use magnesium metal instead of sodium because it forms a more insoluble hydroxide, thereby shifting the equilibrium more completely from metal alkoxide to metal hydroxide. The magnesium should be activated with iodine (or a small amount of methyl iodide), and the water content should be low, otherwise the magnesium will be deactivated. If the amount of water is large, it should be removed by azeotropic distillation (see below), or by drying over anhydrous  $\text{MgSO}_4$  (not  $\text{CaCl}_2$  which combines with alcohols). Acidic materials can be removed by treatment with anhydrous  $\text{Na}_2\text{CO}_3$ , followed by a suitable drying agent, such as calcium hydride, and fractional distillation, using gas chromatography to establish the purity of the product (Ballinger & Long, *J Am Chem Soc* **82** 795 1960). Alternatively, the alcohol can be refluxed with freshly ignited  $\text{CaO}$  for 4 hours and then fractionally distilled (McCurdy & Laidler, *Can J Chem* **41** 1867 1963).

With higher-boiling alcohols it is advantageous to add some freshly prepared magnesium ethoxide solution (only slightly more than required to remove the water), followed by fractional distillation. Alternatively, in such cases, water can be removed by azeotropic distillation with toluene. Higher-melting alcohols can be purified by crystallisation from methanol or ethanol, toluene/petroleum ether or petroleum ether. Sublimation in vacuum, molecular distillation and gas and liquid chromatographic methods are also useful means of purification. For purification *via* derivatives, *vide supra*.

### Polyhydric alcohols

These alcohols are more soluble in water than are monohydric alcohols. Liquids can be freed from water by shaking with type 4A Linde molecular sieves and can safely be distilled only under high vacuum. Carbohydrate alcohols can be crystallised from strong aqueous solution or, preferably, from mixed solvents such as ethanol/petroleum ether or dimethyl formamide/toluene. Crystallisation usually requires seeding and is extremely slow. Further purification can be effected by conversion to the acetyl or benzoyl derivatives which are much less soluble in water and which can readily be recrystallised, e.g. from ethanol. Hydrolysis of the acetyl derivatives, followed by removal of acetate or benzoate and metal ions by ion-exchange chromatography, gives the purified material. On no account should solutions of carbohydrates be concentrated above 40° because of darkening and formation of *caramel*. Ion exchange, charcoal or cellulose column chromatography has been used for the purification and separation of carbohydrates.

### ALDEHYDES

Common impurities found in aldehydes are the corresponding alcohols, aldols and water from self-condensation, and the corresponding acids formed by autoxidation. Acids can be removed by shaking with aqueous 10% sodium bicarbonate solution. The organic liquid is then washed with water. It is dried with anhydrous sodium sulfate or magnesium sulfate and then fractionally distilled. Water soluble aldehydes must be dissolved in a suitable solvent such as diethyl ether before being washed in this way. Further purification can be effected *via* the bisulfite derivative (see above) or the Schiff base formed with aniline or benzidine. Solid aldehydes can be dissolved in diethyl ether and purified as above. Alternatively, they can be steam distilled, then sublimed and crystallised from toluene or petroleum ether.

### AMIDES

Amides are stable compounds. The lower-melting members (such as acetamide) can be readily purified by fractional distillation. Most amides are solids which have low solubilities in water. They can be recrystallised from large quantities of water, ethanol, ethanol/ether, aqueous ethanol, chloroform/toluene, chloroform or acetic acid. The likely impurities are the parent acids or the alkyl esters from which they have been made. The former can be removed by thorough washing with aqueous ammonia followed by recrystallisation, whereas elimination of the latter is by trituration or recrystallisation from an organic solvent. Amides can be freed from solvent or water by drying below their melting points. These purifications can also be used for sulfonamides and acid hydrazides.

### AMINES

The common impurities found in amines are nitro compounds (if prepared by reduction), the corresponding halides (if prepared from them) and the corresponding carbamate salts. Amines are dissolved in aqueous acid, the pH of the solution being at least three units below the  $pK_a$  value of the base to ensure almost complete formation of the cation. They are extracted with diethyl ether to remove neutral impurities and to decompose the carbamate salts. The solution is then made strongly alkaline and the amines that separate are extracted into a suitable solvent (ether or toluene) or steam distilled. The latter process removes coloured impurities. Note that chloroform cannot be used as a solvent for primary amines because, in the presence of alkali, poisonous carbylamines (isocyanides) are formed. However, chloroform is a useful solvent for the extraction of heterocyclic bases. In this case it has the added advantage that while the extract is being freed from the chloroform most of the moisture is removed with the solvent.

Alternatively, the amine may be dissolved in a suitable solvent (e.g. toluene), and dry HCl gas is passed through the solution to precipitate the amine hydrochloride. This is purified by recrystallisation from a suitable solvent mixture (e.g. ethanol/diethyl ether). The free amine can be regenerated by adding sodium hydroxide and isolated as above. Cartridges and columns are available (e.g. KP-NH silica column with slightly nitrogenous alkaline chemistry developed by Biotage Inc (<[www.biotage.com](http://www.biotage.com)>)) for the purification of amines using flash chromatography.

Liquid amines can be further purified *via* their acetyl or benzoyl derivatives (*vide supra*). Solid amines can be recrystallised from water, alcohol, toluene or toluene-petroleum ether. *Care should be taken in handling large quantities of amines because their vapours are harmful (possibly carcinogenic) and they are readily absorbed through the skin.*

### AMINO ACIDS

Because of their zwitterionic nature, amino acids are generally soluble in water. Their solubility in organic solvents rises as the fat-soluble portion of the molecule increases. The likeliest impurities are traces of salts, heavy metal ions, proteins and other amino acids. Purification of these is usually easy, by recrystallisation from water or ethanol/water mixtures. The amino acid is dissolved in the boiling solvent, decolourised if necessary by boiling with 1g of acid-washed charcoal/100g amino acid, then filtered hot, chilled, and set aside for several hours to crystallise. The crystals are filtered off, washed with ethanol, then ether, and dried.

Amino acids have high melting or decomposition points and are best examined for purity by paper or thin layer chromatography. The spots are developed with ninhydrin. Customary methods for the purification of small quantities of amino acids obtained from natural sources (i.e. 1–5g) are ion-exchange chromatography (see Chapter 1). For general treatment of amino acids see Greenstein and Winitz [*The Amino Acids*, Vols 1–3, J.Wiley & Sons, New York 1961] and individual amino acids in Chapters 4 and 7.

A useful source of details such as likely impurities, stability and tests for homogeneity of amino acids is *Specifications and Criteria for Biochemical Compounds*, 3rd edn, National Academy of Sciences, USA, 1972.

### ANHYDRIDES

The corresponding acids, resulting from hydrolysis, are the most likely impurities. Distillation from phosphorus pentoxide, followed by fractional distillation, is usually satisfactory. With high boiling or solid anhydrides, another method involves boiling under reflux for 0.5–1 hours with acetic anhydride, followed by fractional distillation. Acetic acid distils first, then acetic anhydride and finally the desired anhydride. Where the anhydride is a solid, removal of acetic acid and acetic anhydride at atmospheric pressure is followed by heating under vacuum. The solid anhydride is then either crystallised as for acid chlorides or (in some cases) sublimed in a vacuum. A preliminary purification when large quantities of acid are present in a solid anhydride (such as phthalic anhydride) is by preferential solvent extraction of the (usually) more soluble anhydride from the acid (e.g. with  $\text{CHCl}_3$  in the case of phthalic anhydride). *All operations with liquid anhydrides should be carried out in a fume cupboard because of their LACHRYMATORY properties. Almost all anhydrides attack skin.*

### CAROTENOIDS

These usually are decomposed by light, air and solvents, so that degradation products are probable impurities. Chromatography and adsorption spectra permit the ready detection of coloured impurities, and separations are possible using solvent distribution, chromatography or crystallisation. Thus, in partition between immiscible solvents, xanthophyll remains in 90% methanol while carotenes pass into the petroleum ether phase. For small amounts of material, thin-layer or paper chromatography may be used, while column chromatography is suitable for larger amounts. Colourless impurities may be detected by IR, NMR or mass spectrometry. The more common separation procedures are described by P. Karrer and E. Jucker in *Carotenoids*, E.A. Braude (translator), Elsevier, NY, 1950.

Purity can be checked by chromatography (on thin-layer plates, Kieselguhr, paper or columns), by UV or NMR procedures. See "Carotenoids" in Chapter 7.

### ESTERS

The most common impurities are the corresponding acid and hydroxy compound (i.e. alcohol or phenol), and water. A liquid ester from a carboxylic acid is washed with 2N sodium carbonate or sodium hydroxide to remove acid material, then shaken with calcium chloride to remove ethyl or methyl alcohols (if it is a methyl or ethyl ester). It is dried with potassium carbonate or magnesium sulfate, and distilled. Fractional distillation then removes residual traces of hydroxy compounds. This method does not apply to esters of inorganic acids (e.g. dimethyl sulfate) which are more readily hydrolysed in aqueous solution when heat is generated in the neutralisation of the excess acid. In such cases, several fractional distillations, preferably under vacuum, are usually sufficient.

Solid esters are easily crystallisable materials. It is important to note that esters of alcohols must be recrystallised either from non-hydroxylic solvents (e.g. toluene) or from the alcohol from which the ester is derived. Thus methyl esters should be crystallised from methanol or methanol/toluene, but not from ethanol, *n*-butanol or other alcohols, in order to avoid alcohol exchange and contamination of the ester with a second ester. Useful solvents for crystallisation are the corresponding alcohols or aqueous alcohols, toluene, toluene/petroleum ether, and chloroform (ethanol-free)/toluene. Esters of carboxylic acids derived from phenols are more difficult to hydrolyse and exchange, hence any alcoholic solvent can be used freely. Sulfonic acid esters of phenols are even more resistant to hydrolysis: they can safely be crystallised not only from the above solvents but also from acetic acid, aqueous acetic acid or boiling *n*-butanol. Note that sulfonic esters of lower alcohols, e.g. methanol, are good alkylating agents.

Fully esterified phosphoric acid and phosphonic acids differ only in detail from the above mentioned esters. Their major contaminants are alcohols or phenols, phosphoric or phosphonic acids (from hydrolysis), and (occasionally) basic material, such as pyridine, which is used in their preparation. Water-insoluble esters are washed thoroughly and successively with dilute acid (e.g. 0.2N sulfuric acid), water, 0.2N sodium hydroxide and water. After drying with calcium chloride they are fractionally distilled. Water-soluble esters should first be dissolved in a suitable organic solvent and, in the washing process, water should be replaced by saturated aqueous sodium chloride. Some esters (e.g. phosphate and phosphonate esters) can be further purified through their uranyl adducts (*vide supra*). Traces of water or hydroxy compounds can be removed by percolation through, or shaking with, activated alumina (about 100g/L of liquid solution), followed by filtration and fractional distillation in a vacuum. For high molecular weight esters (which cannot be distilled without some decomposition) it is advisable to carry out distillation at as low a pressure as possible. Solid esters can be crystallised from toluene or petroleum ether. Alcohols can be used for recrystallising phosphoric or phosphonic esters of phenols.

### ETHERS

The purification of diethyl ether (see Chapter 4) is typical of liquid ethers. The most common contaminants are the alcohols or hydroxy compounds from which the ethers are prepared, their oxidation products (e.g. aldehydes), peroxides and water. Dialkyl ethers form peroxides much more readily than other ethers, e.g. ethyl phenyl ethers, on standing in air. Peroxides, aldehydes and alcohols can be removed by shaking with alkaline potassium permanganate solution for several hours, followed by washing with water, concentrated sulfuric acid [CARE], then water. After drying with calcium chloride, the ether is distilled. It is then dried with sodium or with lithium aluminium hydride, redistilled and given a final fractional distillation. The drying process is repeated if necessary.

Alternatively, methods for removing peroxides include leaving the ether to stand in contact with iron filings or copper powder, shaking with a solution of ferrous sulfate acidified with N sulfuric acid, shaking with a copper-zinc couple, passage through a column of activated alumina, and refluxing with phenothiazine. Cerium(III) hydroxide has also been used.

A simple test for ether peroxides is to add 10ml of the ether to a stoppered cylinder containing 1ml of freshly prepared 10% solution of potassium iodide containing a drop of starch indicator. No colour should develop during one minute if free from peroxides. Alternatively, a 1% solution of ferrous ammonium sulfate, 0.1M in sulfuric acid and 0.01M in potassium thiocyanate should not increase appreciably in red colour when shaken with two volumes of the ether. Merck-Chemicals supply peroxide test kits (Perex Test) which use a colorimetric method with test strips which can be used to estimate the amount of hydrogen peroxide, from as low a concentration as 0.2mg/L to as high as 1000mg/L. They are very convenient as they can give an indication of the concentration of peroxide rapidly <see <http://www.merck-chemicals.com>>. As a safety precaution against **EXPLOSION** in case purification from peroxides has been insufficiently thorough, at least a quarter of the total volume of liquid ether should remain in the distilling flask when the distillation is discontinued, as the peroxides are generally higher boiling than the corresponding ethers. To minimise peroxide formation, ethers should be stored in dark bottles and, if they are liquids, they should be left in contact with type 4A Linde molecular sieves, in a cold place, over sodium amalgam. The rate of formation of peroxides depends on storage conditions and is accelerated by metal impurities, heat, light, air and moisture. Always be vigilant and test for peroxides. The formation of peroxides is inhibited in the presence of diphenylamine, di-*tert*-butylphenol, or other antioxidants which can be used as stabilisers.

Ethers that are solids (e.g. phenyl ethers) can be steam distilled from an alkaline solution which will hold back any phenolic impurity. After the distillate is made alkaline with sodium carbonate, the insoluble ether is collected either by extraction (e.g. with chloroform, diethyl ether or toluene) or by filtration. It is then crystallised from alcohols, alcohol/petroleum ether, petroleum ether, toluene or mixtures of these solvents, sublimed in a vacuum and recrystallised if necessary.

### HALIDES

Aliphatic halides are likely to be contaminated with halogen acids and the alcohols from which they have been prepared, whereas in aromatic halides the impurities are usually aromatic hydrocarbons, amines or phenols. In both groups the halogen is less reactive than it is in acid halides. Purification is by shaking with concentrated hydrochloric acid, followed by washing successively with water, 5% sodium carbonate or bicarbonate, and water. After drying with calcium chloride, the halide is distilled and then fractionally distilled using an efficient column. For a solid halide the above purification is carried out by dissolving it in a suitable solvent such as toluene. Solid halides can also be purified by chromatography using an alumina column and eluting with toluene or petroleum ether.

They can be crystallised from toluene, petroleum ethers, toluene/petroleum ether or toluene/chloroform/petroleum ether. Care should be taken when handling organic halogen compounds because of their **HIGH TOXICITY**. It should be noted that methyl iodide is a cancer suspect.

Liquid aliphatic halides are obtained alcohol-free by distillation from phosphorus pentoxide. They are stored in dark bottles to prevent oxidation and, in some cases, the formation of phosgene.

A general method for purifying *chlorohydrocarbons* uses repeated shaking with concentrated sulfuric acid [CARE] until no further colour develops in the acid, then washing with water followed by a solution of sodium bicarbonate, then with water again. After drying with calcium chloride, the chlorohydrocarbon is fractionally redistilled to constant boiling point or recrystallised.

## HYDROCARBONS

Gaseous hydrocarbons are best freed from water and gaseous impurities by passage through suitable adsorbents and (if olefinic material is to be removed) oxidants such as alkaline potassium permanganate solution, followed by fractional cooling (see Chapter 1 for cooling baths) and fractional distillation at low temperature. To effect these purifications and also to store the gaseous sample, a vacuum line is necessary.

Impurities in hydrocarbons can be characterised and evaluated by gas chromatography and mass spectrometry. The total amount of impurities present can be estimated from the thermometric freezing curve.

Liquid aliphatic hydrocarbons are freed from aromatic impurities by shaking with concentrated sulfuric acid [CARE] whereby the aromatic compounds are sulfonated. Shaking is carried out until the sulfuric acid layer remains colourless for several hours. The hydrocarbon is then freed from the sulfuric acid and the sulfonic acids by separating the two phases and washing the organic layer successively with water, 2N sodium hydroxide, and water. It is dried with  $\text{CaCl}_2$  or  $\text{Na}_2\text{SO}_4$ , and then distilled. The distillate is dried with sodium wire,  $\text{P}_2\text{O}_5$ , or metallic hydrides, or passage through a dry silica gel column, or preferably, and more safely, with molecular sieves (see Chapter 1) before being finally fractionally distilled through an efficient column. If the hydrocarbon is contaminated with olefinic impurities, shaking with aqueous alkaline permanganate is necessary prior to the above purification. Alicyclic and paraffinic hydrocarbons can be freed from water, non-hydrocarbon and aromatic impurities by passage through a silica gel column before the final fractional distillation. This may also remove isomers. (For the use of chromatographic methods to separate mixtures of aromatic, paraffinic and alicyclic hydrocarbons see references in the bibliography in Chapter 1 under *Liquid and Flash Chromatography, Gas Chromatography and High Performance Liquid and Flash Chromatography*). Another method of removing branched-chain and unsaturated hydrocarbons from straight-chain hydrocarbons depends on the much faster reaction of the former with chlorosulfonic acid.

Isomeric materials which have closely similar physical properties can be serious contaminants in hydrocarbons. With aromatic hydrocarbons, e.g. xylenes and alkyl benzenes, advantage is taken of differences in ease of sulfonation. If the required compound is sulfonated more readily, the sulfonic acid is isolated, crystallised (e.g. from water), and decomposed by passing superheated steam through the flask containing the acid. The sulfonic acid undergoes hydrolysis, and the liberated hydrocarbon distils with the steam. It is separated from the distillate, dried, distilled and then fractionally distilled. For small quantities (10–100mg), vapour phase chromatography is the most satisfactory method for obtaining a pure sample (for column materials for packings see Chapter 1).

Azeotropic distillation with methanol or 2-ethoxyethanol (cellosolve) has been used to obtain highly purified saturated hydrocarbons and aromatic hydrocarbons such as xylenes and isopropylbenzenes.

Carbonyl-containing impurities can be removed from hydrocarbons (and other oxygen-lacking solvents such as  $\text{CHCl}_3$  and  $\text{CCl}_4$ ) by passage through a column of Celite 545 (100g) mixed with concentrated sulfuric acid (60ml). After first adding some solvent and about 10g of granular  $\text{Na}_2\text{SO}_4$ , the column is packed with the mixture and a final 7-8cm of  $\text{Na}_2\text{SO}_4$  is added at the top [Hornstein & Crowe, *Anal Chem* **34** 1037 1962]. Alternatively, Celite impregnated with 2,4-dinitrophenylhydrazine can be used.

With solid hydrocarbons such as naphthalene and polycyclic hydrocarbons, preliminary purification by sublimation in vacuum (or high vacuum if the substance is high melting) is followed by zone refining and finally by chromatography (e.g. on alumina) using low-boiling liquid hydrocarbon eluents. These solids can be recrystallised from alcohols, alcohol/petroleum ether or from liquid hydrocarbons (e.g. toluene) and dried below their melting points. Aromatic hydrocarbons that have been purified by zone melting include anthracene, biphenyl, fluoranthrene, naphthalene, perylene, phenanthrene, pyrene and terphenyl, among others. Some polycyclic hydrocarbons, e.g. benzopyrene, are **CARCINOGENIC**.

Olefinic hydrocarbons have a very strong tendency to polymerise, and commercially available materials are generally stabilised, e.g. with hydroquinone. When distilling compounds such as vinylpyridine or styrene, the stabiliser remains behind and the purified olefinic material is more prone to polymerisation. The most common impurities are higher-boiling dimeric or polymeric compounds. Vacuum distillation in a nitrogen

atmosphere not only separates monomeric from polymeric materials but in some cases also depolymerises the impurities. The distillation flask should be charged with a polymerisation inhibitor, and the purified material should be used immediately or stored in the dark and mixed with a small amount of stabiliser (e.g. 0.1% of hydroquinone or di-*tert*-butylcatechol). It is also advisable to add to the flask a small amount (*ca* 5-10% by volume of liquid in the flask) of a ground mixture of Kieselguhr and NaCl which will provide nuclei for facilitating boiling and finally for cleaning the flask from insoluble polymeric residue (due to the presence of the water soluble NaCl).

### IMIDES

Imides (e.g. phthalimide) can be purified by conversion to their potassium salts by reaction in ethanol with ethanolic potassium hydroxide. The imides are regenerated when the salts are hydrolysed with water or dilute acid. Like amides, imides readily crystallise from alcohols and, in some cases (e.g. quinolinic imide), from glacial acetic acid.

### IMINO COMPOUNDS

These substances contain the  $-C=NH$  group and, because they are strong, unstable bases, they are kept as their more stable salts, such as the hydrochlorides. (The free base usually hydrolyses to the corresponding oxo compound and ammonia.) Like amine hydrochlorides, the salts are purified by solution in alcohol containing a few drops of hydrochloric acid. After treatment with charcoal, and filtering, dry diethyl ether (or petroleum ether if ethanol is used) is added until crystallisation sets in. The salts are dried and kept in a vacuum desiccator.

### KETONES

Ketones are more stable to oxidation than aldehydes and can be purified from oxidisable impurities by refluxing with potassium permanganate until the colour persists, followed by shaking with sodium carbonate (to remove acidic impurities) and distilling. Traces of water can be removed with type 4A Linde molecular sieves. Ketones which are solids can be purified by crystallisation from alcohol, toluene, or petroleum ether, and are usually sufficiently volatile for sublimation in vacuum. Ketones can be further purified *via* their bisulfite, semicarbazone or oxime derivatives (*vide supra*). The bisulfite addition compounds are formed only by aldehydes and methyl ketones but they are readily hydrolysed in dilute acid or alkali.

**MACROMOLECULES** See Chapter 7.

### NITRILES

*All purifications should be carried out in an efficient fume cupboard because of the TOXIC nature of these compounds.*

Nitriles are usually prepared either by reacting the corresponding halide or diazonium salts with a cyanide salt or by dehydrating an amide. Hence, possible contaminants are the respective halide or alcohol (from hydrolysis), phenolic compounds, amines or amides. Small quantities of phenols can be removed by chromatography on alumina. More commonly, purification of liquid nitriles or solutions of solid nitriles in a solvent such as diethyl ether is by shaking with dilute aqueous sodium hydroxide, followed by washing successively with water, dilute acid and water. After drying with sodium sulfate, the solvent is distilled off. Liquid nitriles are best distilled from a small amount of  $P_2O_5$  which, besides removing water, dehydrates any amide impurity to the nitrile. About one-fifth of the nitrile should remain in the distilling flask at the end of the distillation (*the residue may contain some inorganic cyanide*). This purification also removes alcohols and phenols. Solid nitriles can be recrystallised from ethanol, toluene or petroleum ether, or a mixture of these solvents. They can also be sublimed under vacuum. Preliminary purification by steam distillation is usually possible.

Strong alkali or heating with dilute acids may lead to hydrolysis of the nitrile and should be avoided.

### NITRO COMPOUNDS

Aliphatic nitro compounds are generally acidic. They are freed from alcohols or alkyl halides by standing for a day with concentrated sulfuric acid, then washed with water, dried with magnesium sulfate followed by calcium sulfate and distilled. The principal impurities are isomeric or homologous nitro compounds. In cases where the nitro compound was originally prepared by vapour phase nitration of the aliphatic hydrocarbon, fractional distillation should separate the nitro compound from the corresponding hydrocarbon. Fractional crystallisation is more effective than fractional distillation if the melting point of the compound is not too low.

The impurities present in aromatic nitro compounds depend on the aromatic portion of the molecule. Thus, benzene, phenols or anilines are probable impurities in nitrobenzene, nitrophenols and nitroanilines, respectively.

Purification should be carried out accordingly. Isomeric compounds are likely to remain as impurities after the preliminary purifications to remove basic and acidic contaminants. For example, *o*-nitrophenol may be found in samples of *p*-nitrophenol. Usually, the *o*-nitro compounds are more steam volatile than the *p*-nitro isomers and can be separated in this way. Polynitro impurities in mononitro compounds can be readily removed because of their relatively lower solubilities in solvents. With acidic or basic nitro compounds which cannot be separated in the above manner, advantage may be taken of their differences in pK values (see Chapter 1). The compounds can thus be purified by preliminary extractions with several sets of aqueous buffers of known pH (see for example Table 24, Chapter 1) from a solution of the substance in a suitable solvent such as diethyl ether. This method is more satisfactory and less laborious the larger the difference between the pK value of the impurity and the desired compound. Heterocyclic nitro compounds require similar treatment to the nitroanilines. Neutral nitro compounds can be steam distilled.

**NUCLEIC ACIDS** See Chapter 7.

### PHENOLS

Because phenols are weak acids, they can be freed from neutral impurities by dissolution in aqueous N sodium hydroxide (cf pK) and extraction with a solvent such as diethyl ether, or by steam distillation to remove the non-acidic material. The phenol is recovered by acidification of the aqueous phase with 2N sulfuric acid, and either extracted with ether or steam distilled. In the second case the phenol is extracted from the steam distillate after saturating it with sodium chloride (salting out). A solvent is necessary when large quantities of liquid phenols are purified. The phenol is fractionated by distillation under reduced pressure, preferably in an atmosphere of nitrogen to minimise oxidation. Solid phenols can be crystallised from toluene, petroleum ether or a mixture of these solvents, and can be sublimed under vacuum. Purification can also be effected by fractional crystallisation or zone refining. For further purification of phenols *via* their acetyl or benzoyl derivatives *vide supra*.

**POLYPEPTIDES AND PROTEINS** See Chapter 7.

### QUINONES

These are neutral compounds which are usually coloured. They can be separated from acidic or basic impurities by extraction of their solutions in organic solvents with aqueous basic or acidic solutions, respectively. Their colour is a useful property in their purification by chromatography through an alumina column with, e.g. toluene, as eluent. They are volatile enough for vacuum sublimation, although with high-melting quinones a very high vacuum is necessary. *p*-Quinones are stable compounds and can be recrystallised from water, ethanol, aqueous ethanol, toluene, petroleum ether or glacial acetic acid. *o*-Quinones, on the other hand, are readily oxidised. They should be handled in an inert atmosphere, preferably in the absence of light.

### SALTS

#### With metal ions

Water-soluble salts are best purified by preparing a concentrated aqueous solution to which, after decolourising with charcoal and filtering, ethanol or acetone is added so that the salts crystallise. They are collected, washed with aqueous ethanol or aqueous acetone, and dried. In some cases, water-soluble salts can be recrystallised satisfactorily from alcohols. With very water-soluble salts, pure crystals are best obtained by dissolving them in water and allowing the solution to evaporate slowly in a desiccator over a suitable desiccant in a cold room. When crystals are formed they are removed, e.g. by centrifugation, washed with a little ice-cold water and dried in a vacuum. Water-insoluble salts are purified by Soxhlet extraction, first with organic solvents and then with water, to remove soluble contaminants. The purified salt is recovered from the thimble.

#### With organic cations

Organic salts (e.g. trimethylammonium benzoate) are usually purified by recrystallisation from polar solvents (e.g. water, ethanol or dimethyl formamide). If the salt is too soluble in a polar solvent, its concentrated solution should be treated dropwise with a miscible non-polar, or less polar, solvent (see Table 8, Chapter 1) until crystallisation begins.

**With sodium alkane sulfonates**

These are purified from sulfites by boiling with aqueous HBr. They are purified from sulfates by adding BaBr<sub>2</sub>. Sodium alkane disulfonates are finally precipitated by addition of MeOH (Pethybridge & Taba *J Chem Soc Faraday Trans 1* **78** 1331 1982).

**SULFUR COMPOUNDS****Disulfides**

These can be purified by extracting acidic and basic impurities with dilute aqueous base or acid, respectively. However, they are somewhat sensitive to strong alkali which slowly cleaves the disulfide bond. The lower-melting members can be fractionally distilled under vacuum. The high members can be recrystallised from alcohol, toluene or glacial acetic acid.

**Sulfones**

Sulfones are neutral and very stable compounds that can be distilled without decomposition. They are freed from acidic and basic impurities in the same way as disulfides. The low-molecular-weight members are quite soluble in water, but the higher members can be recrystallised from water, ethanol, aqueous ethanol or glacial acetic acid.

**Sulfoxides**

These are odourless, rather unstable compounds because they oxidise to sulfones, and should be distilled under vacuum in an inert atmosphere. They are generally water-soluble but can be extracted from aqueous solution with a solvent such as diethyl ether.

**Thioethers (sulfides)**

Thioethers are neutral stable compounds that can be freed from acidic and basic impurities as described for disulfides. They can be recrystallised from organic solvents and distilled without decomposition. They have sulfurous odours.

**Thiols**

Thiols, or mercaptans, are stronger acids than the corresponding aliphatic hydroxy or phenolic compounds, but can be purified in a similar manner. However, care must be exercised in handling thiols to avoid their oxidation to disulfides. For this reason, purification is best carried out in an inert atmosphere in the absence of oxidising agents. Similarly, thiols should be stored out of contact with air. They can be distilled without change, and the higher-melting thiols (which are usually more stable) can be crystallised, e.g. from water or dilute alcohol. They oxidise readily in alkaline solution but can be separated from the disulfide which is insoluble in this medium. They should be stored in the dark below 0°. *All operations with thiols should be carried out in an efficient fume cupboard because of their very unpleasant odour and their TOXICITY.*

**Thiolsulfonates (disulfoxides)**

Thiolsulfonates are neutral and are somewhat light-sensitive compounds. Their most common impurities are sulfonyl chlorides (neutral) or the sulfonic acid or disulfide from which they are usually derived. The first can be removed by partial freezing or crystallisation, the second by shaking with dilute alkali, and the third by recrystallisation because of the higher solubility of the disulfide in solvents. Thiolsulfonates decompose slowly in dilute, or rapidly in strong, alkali to form disulfides and sulfonic acids. Thiolsulfonates also decompose on distillation but they can be steam distilled. The solid members can be recrystallised from water, alcohols or glacial acetic acid.

**PURIFICATION via FLUOROCHROMES**

If the purification procedure is proving difficult then by tagging the desired molecule in a mixture with a fluorochrome (see Table 20, Chapter 1) can provide a means of following the substance through the purification process. The fluorochrome should have a group which can react with the desired compound, and it should be possible to remove the fluorochromic group after purification. Such groups are present for example in fluorescein-isothiocyanate (FITC), where the SCN group can react with an RNH<sub>2</sub> compound to form fluorescent thioureas; 4-bromomethyl-7-methoxycoumarin, which can react with R-OH or R-COONa compounds to form fluorescent ether or ester links with the desired compound; or 3-aminocoumarin, which reacts through its NH<sub>2</sub> group to form fluorescent amides. The fluorescent products can then be readily identified by their fluorescence, separated from impurities, and it should be possible to recover the purified compound after chemically separating it from the fluorescent tag. Such procedures can also be used as analytical tools for detecting specific substances (see *fluorescence spectra* in Chapter 1 and Table 20).

## BIBLIOGRAPHY

For earlier bibliographies see *Purification of Laboratory Chemicals*, 4th, 5<sup>th</sup>, and 6<sup>th</sup> Edns, ISBN 0750628391 (1996, hardback) and 0750637617 (1997, paperback), 0750675713 (2003), and 9781856175678 (2009). See Bibliography in Chapter 1, and visit the websites of DA Books <www.dadirect.com>, Elsevier <www.elsevierdirect.com> and Wiley <www.wiley.com> to update this bibliography.

## CHARACTERISATION OF ORGANIC AND INORGANIC COMPOUNDS

F. Fiegl, *Spot Tests in Organic Analysis*, 7th Edn, Elsevier Science, New York, 1989. ISBN 0444402098.

*Fieser and Fieser's Reagents for Organic Synthesis* (L. Fieser, M. Fieser and M.B. Smith Eds.), J.Wiley & Sons, Inc., New York, 2007, Volumes 1-23, and collective index 1-23. ISBN 9780470133910.

W.L. Jolly, *The Synthesis and Characterisation of Inorganic Compounds*, Waveland Press, 1991. ISBN 0881335789.

D. Kiemble, R.M. Silverstein and F.X. Webster, *Spectrometric Identification of Organic Compounds*, 7th Edn, J. Wiley & Sons, New York, 2005. ISBN 9780471393627.

R.B. King (Ed.), *Encyclopedia of Inorganic Chemistry* (10 volume set), J. Wiley & Sons, New York, 2005. ISBN 9780470860786.

W.M. Haynes, *CRC Handbook of Chemistry and Physics*, 92nd Edn, CRC Press, 2011. ISBN 9781439855119.

J. Mendham, R.C. Denney, J.D. Barnes and M.J.K. Thomas, *Vogel's Quantitative Chemical Analysis*, 6th Edn, Prentice Hall, Harlow, 2000. ISBN 0582226287.

J.R. Morig, C.N. Hammond and P.F. Schatz, *Techniques in Organic Chemistry*, W.H. Freeman & Co., N.Y., 2010. ISBN 10: 1429219564, 13: 9781429219563.

T.C. Morrill, R.L. Shriner and C.K.F. Hermann, *The Systematic Identification of Organic Compounds*, 8th Edn, J. Wiley & Sons, New York, 2003. ISBN 9780471215035 and 9780471451655.

L.A. Paquette (Ed.), *The Encyclopedia of Reagents for Organic Synthesis*, J. Wiley & Sons, New York, 2009. ISBN 9780470017548.

M.C. Pirrung, *The Synthetic Organic Chemist's Companion*, J.Wiley & Sons Inc, 2007. ISBN 9780470107072.

## METAL HYDRIDES

A. Dedieu, (Ed.), *Transition Metal Hydrides*, VCH-Wiley Publishing, 1992. ISBN 9783527279852.

M.V.C. Sastri (Ed.), *Metal Hydrides: Fundamentals and Applications*, Springer, 1999. ISBN 9783540642046.

J. Seyden-Penne, *Reductions by the Alumino- and Borohydrides in Organic Synthesis*, 2nd Edn., J. Wiley & Sons, New York, 1997. ISBN 9781560819391.

**SPECTROSCOPY**

R.J. Anderson and D. Bendell, *Organic Spectroscopic Analysis*, Royal Society of Chemistry, 2004. ISBN 9780854044764.

G.W. Gokel, *Deane's Handbook of Organic Chemistry*, McGraw-Hill, 2004. ISBN 0071375937.

M. Gunzler and M. Heise, *IR Spectroscopy; An Introduction*, VCH-Wiley Publishing, 2001. ISBN 9783527288960.

F.W. McLafferty and D.B. Stauffer, *The Wiley/NBS Registry of Mass Spectral Data*, (7 volumes), J. Wiley & Sons, New York, 1989. ISBN 0471628867.

Y-C. Ning, *Structural Identification of Organic Compounds with Spectroscopy*, VCH-Wiley Publishing, 2005. ISBN 9783527312405.

C.J. Pouchert, *The Aldrich Library of FT-IR Spectra*, Aldrich Chemical Co., Milwaukee, 3 Volume set, 2nd Edn., 1997. ISBN 0941633349X.

C.J. Pouchert and J. Behnke, *The Aldrich Library of  $^{13}\text{C}$  and  $^1\text{H}$  FT NMR Spectra*, 3 Volume set, Aldrich Chemical Co., Milwaukee, 1993. ISBN 0941633349.

**TRACE METAL ANALYSIS**

T.J.S. Bruno and D.N. Paris, *Handbook of Basic Tables for Chemical Analysis*, CRC Press LLC, 2003. ISBN 9780849315732.

G. Christian and M. Epstein, *Atomic Absorption Spectrometry*, American Chemical Society, 1990. ISBN 9780841210998.

R.Cornelis, J. Caruso and H. Crews, *Handbook of Elemental Speciation: Techniques and Methodology* Vol 1, J. Wiley & Sons, 2003. ISBN 9780471492146.

R. Dean, *Atomic Absorption and Plasma Spectroscopy*, 2nd Edn., J. Wiley & Sons, 1997. ISBN 9780471972549.

I.S. Krull (Ed.), *Trace Metal Analysis and Speciation*, Elsevier Science, 1991. ISBN 9780444882097.

# CHAPTER 3

## THE FUTURE OF PURIFICATION

### INTRODUCTION

The essence of research is to seek answers wherever there are questions. Regardless of what the answers are, the experiments to be conducted must be carried out with utmost care. For this, one must ensure that the quality of the reactants used and the products obtained are of the highest possible purity. In general terms, one can broadly categorise experimental chemistry and biological chemistry into the following areas:

Isolation and identification of substances (natural products from nature, protein purification and characterisation, etc).

Synthesis of substances (organic, or inorganic in nature; these substances may be known substances or new compounds).

Analysis of substances (this is a key process in the identification of new or known chemical and biological substances. Methods of analysis include spectroscopic methods, derivatisation and sequencing methods).

Measurements of particular properties of a compound or substance (enzyme kinetics, reaction kinetics, FACS, fluorescence-activated cell sorting, assay).

Impressive and sophisticated strategies, in the form of new reagents, catalysts and chemical transformations, are currently available for the synthesis of molecules. In recent years there has been a deviation in focus from developing new synthetic routes and reactions to improving methods for carrying out reactions. In particular, traditional reactions are carried out in new ways such that the efficiencies of reactions are greatly improved. Emphasis on “Green Technologies” and “Green Chemistry” is continually being stressed to the experimenter and will be for many years to come. Included in this is the movement towards the development of “greener” technologies. Although there is continual debate over the definition of “green” technologies, for our intent and purpose we consider these to include methodologies that are developed to reduce waste and energy, improve on atom economy and are environmentally benign. [Anastas & Kirchhoff *Acc Chem Res* **35** 686 2002; *Benign by Design*; American Chemical Society: Washington, DC, 1994; Constable et al, *Green Chem* **9** 411 2007; Li & Trost *PNAS* **105** 13197 2008; Anastas & Eghbali, *Chem Soc. Rev.* **39** 301 2010]. This may mean that there will be a new world order for carrying out synthesis, which in turn will change the way that one approaches purification. Hence the future of purification very much depends on future methods and trends in synthesis. In the previous edition of this book, some improved methods of synthesis which seek to minimise purification steps were outlined. These included the use of solid phase synthesis, fluoros chemistry as well as ionic liquids. Although some of these methods are still important in some areas of research, other trends have emerged. In this chapter, a brief survey of the emerging trends is presented.

Safety is another issue that goes hand in hand with “Green Science”, and cannot be emphasised enough. It should be the first consideration when planning a scientific operation. Safety issues have already been discussed in Chapter 1, and should never be considered a *nuisance*, or an aspect which hampers procedures. This is no longer a *local* or a *regional* issue, but is becoming more and more a *global* problem. It is true that it means more effort should be put in planning an operation, but this is something that the worker should get used to. It is totally *unacceptable* today, and certainly will be in the future, not to take safety very seriously. Questions to be addressed include: Could this procedure be carried out in a safer way with respect not only to the operator but also to those around him or her? How does it affect the immediate as well as the extended environment around the operation? Could the operation be performed in an alternative way that can better satisfy these criteria? Safety now goes well beyond the laboratory and all should be consciously aware of the consequences of a potentially unsafe situation.

## ATOM ECONOMY

Atom economic chemistry seeks to maximise the incorporation of starting materials into the final product of any given reaction and thus reduce side products or waste formed in a chemical transformation. [Trost *Angew Chem, Int Ed* **34** 259 2005.] Thus in an ideal atom economic reaction, all the atoms present in the starting materials are present in the product. Ways to achieve atom economy or efficient reactions include the use of catalysts to enable transformations that will otherwise utilise multiple steps or are impossible to perform. Cross coupling reactions to form C-N, C-O, C-C bonds can now be achieved with relative ease. Despite the progress in the area of catalysis, there are continual challenges to reduce the number of steps needed to achieve the desired transformations. For example, direct methods for C-H functionalisation in a selective and predictable fashion can dramatically alter the way organic molecules are assembled, negating multi-step chemical processes and rendering functional group interconversion steps obsolete. To date, there are numerous reported protocols, with new catalyst systems and designs, for the direct transformation of C-H bond to C-C or C-heteroatom bonds [for a sample of articles on C-H functionalisation, see the special issue on C-H functionalisation in organic synthesis themed issue, Eds. Davies, Du Bois & Yu, in *Chem Soc Rev* Issue 4, 2011]. This field is an emerging, highly competitive area of research and a comprehensive review of C-H functionalisation is beyond the scope of this book.

Even a seemingly 'simple' transformation such as amide bond formation warrants investigation of new atom economic methods. This is because 'traditional' methods for amide formation suffer from the drawback of producing a stoichiometric amount of waste. Advances in this area are metal catalysed approaches to amide bond formation that includes oxidative amidation of alcohols [for examples, see Taylor et al. *Synlett* 1293 2002; Owston et al. *Org. Lett.* **9** 73 2007; Zweifel et al. *Angew Chem, Int Ed* **48** 559 2009] and amines [e.g. Kim et al. *Angew Chem, Int Ed* **47** 9249 2008], aminocarbonylation of aryl halides, [for example, Ueda et al. *J Am Chem Soc* **122** 10722 2000; Schnyder, et al. *J Org Chem* **66** 4311 2001; Wu et al. *Chem Asian J* **5** 2168 2010], rearrangement of aldoximes [e.g. Shie & Fang *J Org Chem* **68** 1158 2003; Owston et al. *Org Lett* **9** 3599 2007], and direct amidation of aldehydes [e.g. Ali & Punniyamurthy *Adv Synth Catal* **352** 288 2010; Shie & Fang *J Org Chem* **68** 1158 2003; Yoo & Li *J Am Chem Soc* **128** 13064 2006].

Improved atom efficiency can also be achieved if it was possible to carry out synthesis without the use of protecting groups. The protection-deprotection sequence of functional groups increases the number of steps in the synthesis of target compounds. As such, novel chemistry is needed to overcome this. In elegant work reported by Baran et al. [*Nature* **446** 404 2007], a total synthesis of a natural product was achieved without any protecting groups.

## ORGANOCATALYSIS

One of the major areas of chemical research is in the development of efficient catalysts for carrying out organic transformations. Traditionally this field is dominated by metal mediated catalysis, although increasingly biocatalysis and organocatalysis are gaining prominence. Organocatalysts are defined as catalysts (usually small organic molecules) with low molecular weights (<1000g/mol) where a metal is not part of the active principle. The potential advantages that organocatalysts can offer over metal catalysts include the ease of handling, good stability, low costs and the environmentally benign nature of the former. A major disadvantage of organocatalysis is the limited substrate scope for a particular organocatalyst in a particular transformation. It should be added that the field of organocatalysis was in its relative infancy and stunning progress has been made in the last ten years or so. For example, readily available organic compounds have been found to be capable of catalysing a range of reactions. These include carbon-carbon bond forming reactions (Aldol, Mannich, Diels-Alder, alkylations, cyclopropanations, etc), epoxidation reactions, desymmetrisation reactions and so on [for an excellent book, see Berkessel and Gröger, in the bibliography]. Classes of compounds that have demonstrated applications as organocatalysts include  $\alpha$ -amino acids [Jarvo & Miller *Tetrahedron* **58** 2481 2002], *N*-heterocyclic carbenes [Marion et al. *Angew Chem, Int Ed* **46** 2988 2007], thioureas [Connon *Chemistry-A Eur J* **12** 5419 2006], cinchona alkaloids [Dalaigh *Syn Lett* 875 2005] and so on. One of the earliest practical demonstrations of the use of organocatalysts is exemplified by the Hajos-Parrish-Eder-Sauer-Wiechert reaction which was patented as early as 1971 [Eder, Sauer & Wiechert DE 2014757 1971, Eder, Sauer & Wiechert *Angew Chem, Int Ed* **10** 496 1971, Hajos & Parrish DE 2102623 1971, Hajos & Parrish *J Org Chem* **39** 1612 1974]. The Robinson aldol annulation reaction utilised the humble proline molecule as the organocatalyst, and is the *first* example of a highly enantioselective carbon-carbon bond forming reaction. Recently, advances in organocatalysis have progressed to the level of sophistication that has enabled the synthesis of relatively complex organic molecules with multiple stereocentres in high chemical and optical yields. There are many excellent reviews and monographs on organocatalysis, and some of these are listed in the bibliography. Also see ChemFiles, Vol 6, no 4 from Sigma-Aldrich for a compilation of some commercially

available organocatalysts. Another class of organocatalysts are the Phase Transfer Catalysts (PTCs) which have been used for quite some time, and chiral PTCs have more recently been used to catalyse reactions stereospecifically. Sections on “Organocatalysts” and “Phase Transfer Catalysts” are presented in Chapter 6, Part 1 and Part 2, respectively.

## MICROWAVE TECHNOLOGIES

The first reports of the use of microwave technologies to accelerate organic reactions were published in 1986 [Gedye et al. *Tetrahedron Lett* **27** 279 1986, Giguere et al. *Tetrahedron Lett.* **27** 4945 1986]. These days, dedicated microwave instrumentation for carrying out organic synthesis is an almost indispensable piece of equipment in chemical laboratories. Microwave assisted organic synthesis have several advantages over conventional heating methods in that the reaction times are greatly reduced, yields are increased, side reactions are decreased and reactions have greater reproducibility. Microwave irradiation produces efficient internal heating due to the direct coupling of the microwave energy with the reactants and solvents. Inverted temperature gradients in microwave *versus* traditional oil-bath heating have been demonstrated. Although there is still some controversy regarding the “special” effects of microwave assisted technologies, it is generally accepted that microwave technologies provide “specific microwave effects” – that resulting due to thermal effects. The microwave thermal effect may be different to thermal effects from conventional heating methods in view of the inverted temperature gradient leading to direct “in-core” heating, or superheating effects of solvents at atmospheric pressure, or selective heating by strongly absorbing components of some of the reactants. Microwave-assisted organic synthesis has been demonstrated in a number of reactions: from transition metal catalysed reactions to solid phase synthesis to polymerisation reactions. For online resources on microwave chemistry, go to <http://www.milestonesci.com> (and follow the links to the resource library). Some excellent reviews and books on this topic are listed in the bibliography. It should be noted that microwave technologies have been touted as an important development towards “green” technologies in terms of reduction in energy and waste as well as improved efficiencies. In many instances, solventless reactions under microwave conditions have also been developed and reported [see e.g. Varma *Green Chem* **1** 43 1999].

There are a number of commercial microwaves for carrying out synthesis. Mono-mode (single mode) reactors direct electromagnetic irradiation onto a reaction vessel which is mounted at a fixed position. In these reactors, only one vessel can be used at any one time. In contrast, a multi-mode reactor can accommodate several reaction vessels which can be irradiated simultaneously. The choice of solvent for microwave reactions is dependent on the loss factor “ $\tan \delta$ ”. A reaction medium with a high “ $\tan \delta$ ” is required for efficient absorption of electromagnetic energy from the microwave and hence leads to rapid heating. High microwave absorbing solvents include ethanol, DMSO, methanol, 1-butanol (“ $\tan \delta$ ”  $>0.5$ ), while medium solvent absorbers (“ $\tan \delta$ ” between 0.1–0.5) include water and DMF. Poor absorbers are solvents like hexane, dichloromethane, chloroform, THF, acetone and ethyl acetate. Common solvents without a permanent dipole, e.g. carbon tetrachloride and benzene are more or less microwave transparent. Solvents with low “ $\tan \delta$ ” values can still be used in microwave reactions as the other components present are likely to be polar and can allow enough heating by microwaves. Alternatively, practices such as doping the reaction with solvents with high “ $\tan \delta$ ” values have also been successful. Of note, ionic liquids (which have high “ $\tan \delta$ ” values) have been used successfully [see Leadbeater, Torenus & Tye *Comb Chem High Throughput Screening* **7** 511 2004, for a review].

## FLOW CHEMISTRY

The traditional way of carrying out synthesis utilises batch chemistry, with chemistry occurring inside flasks or reaction vessels. Although there are limitations with batch chemistry, especially in the scale-up of reactions, this has been the mainstay of organic synthesis. In recent years, continuous flow chemistry, an old process concept, has gained increasing attention as an alternative technology to batch chemistry. This will clearly have an impact on the way we carry out synthesis or think about synthesis. The advantages of continuous flow chemistry are improved safety and process reliability, reproducibility and facile automation. Continuous flow chemistry can be carried out on different scales, from using microfluidics (miniaturisation) to large continuous reactors found in chemical plants. In a research laboratory setting, there are advantages and disadvantages in using microfluidics *versus* minifluidic reactors as is summarised by Kirschning et al. [*Chem Commun* **47** 4583 2011]. There are a number of vendors that offer flow chemistry equipment, with varying capabilities and price ranges [e.g. <http://www.vapourtec.co.uk/>; <http://www.syrris.com/>]. One of the hurdles to the ready adoption of flow chemistry in a research laboratory setting is the need to consider physical chemistry and engineering principles such as mixing times, residence time, flow rates, etc. when designing an experiment. This can initially be a daunting task to a traditional organic chemist. One of the early pioneers of flow chemistry in organic

synthesis is S.V. Ley, University of Cambridge, UK. The reader is advised to refer to his numerous, elegant work in this area – from the development of solid supported reagents and scavengers to applications to complex natural product synthesis. Many types of reactions can be carried out using flow chemistry, and flow chemistry has particular advantage when carrying out “dangerous” reactions such as oxidations. Also see Aldrich Chem Files, Vol 5, no 7 for enabling technologies: Microreactor Technology.

## SOLID PHASE SYNTHESIS

The promise of cleaner, more rapid and efficient chemistries *via* solid phase synthesis (SPS) was the driving force behind the huge body of research that emerged. The ease of work-up and purification procedures in solid phase as compared to solution phase chemistry, as well as the scope for combinatorial chemistry provided impetus for further development in this field. The earliest studies on solid phase chemistry were focused on solid phase peptide synthesis (SPPS). The concept of carrying out reactions on a polymer support as distinct to reactants in solution, was conceived by R.B. Merrifield who received the Nobel Prize in Chemistry in 1984 for his pioneering work. However, since the mid-1990s, advances in solid phase chemistry have moved beyond the routine (often robotic) synthesis of small to medium sized peptides and oligonucleotides. SPOS (solid phase organic synthesis) gained much prominence due to the wealth of compounds (combinatorial libraries) that can be synthesised rapidly. This is especially important for pharmaceutical companies screening for compounds with certain biological profiles, or for chemical companies screening for new catalysts or reagents. However, the lack of generality of reactions carried out on solid support as well as the difficulties in monitoring reactions on solid phase has limited the synthetic applications of SPOS. The legacy from solid phase synthesis that is more widely adopted in recent times is the use of solid supported reactants and reagents for carrying out synthesis and for scavenging and purification purposes. The more common methods of SPS are outlined here. For a more complete treatise, readers are encouraged to consult the 5th and 6th editions of this book.

### SOLID PHASE PEPTIDE SYNTHESIS (SPPS)

Extensive studies on the synthesis of peptides on solid phase have been carried out, so much so that the technique of SPPS can be reliably and routinely used for the synthesis of short peptides by novices in the field. A large number of resins and reagents have been developed specifically for this purpose, and much is known about problems and avoidance of racemisation, difficult couplings, compatibility of reagents and solvents. Methods for monitoring the success of coupling reactions are available. Automated synthesisers are available commercially (e.g. from Protein Technologies, Rainin Inst Inc, Tuscon AZ; See Google) which can carry out as many as a dozen polypeptide syntheses simultaneously. A more recent advance is in the use of microwave energy to carry out the reactions (e.g. Liberty™ Microwave Peptide Synthesis, cf <http://www.cem.com>). By using different solid supports, protected amino acids and slightly different chemistry but the same equipment, peptide synthesis can be accomplished from the carboxy or the amino terminal with equal success [refer to CEM Corporation catalogues]. The most satisfactory chemistry currently used is Fmoc (9-fluorenylmethoxycarbonyl) chemistry whereby the amino group of the individual amino acid residues is protected as the Fmoc. A large number of Fmoc-amino acids are commercially available as well as polymer resins to which the specific Fmoc-amino acid (which will eventually become the carboxy terminal residue of the peptide) is attached. With automated synthesisers, the solvent used is *N*-methylpyrrolidone and washings are carried out with dimethylformamide. A cycle for one residue varies with the residue but can take an hour or more. This means that 70–80 mer polypeptides could take more than a week to prepare. This is not a serious drawback because several different polypeptides can be synthesised simultaneously. The success of the synthesis is dependent on the amino acid sequence since there are some twenty or more different amino acids and the facility of forming a peptide bond varies with the pair of residues involved. However, generally 70 to 80 mers are routinely prepared, and if the sequence is favourable, up to 120-mer polypeptides can be synthesised. After deprotection, the polypeptide is usually purified by HPLC using a C18 column with reverse phase chromatography. A new paradigm in solid phase peptide synthesis developed by A.G. Livingston and co-workers [So et al. *Organic Process Research and Development* **14** 1313 2010] is the use of solvent resistant nano filters (SRNFs) which adopt the newly emerging technology of organic solvent nanofiltration (OSN). Thus after the activation and coupling reactions, reagents and unreacted compounds are removed by nanofiltration, and then again after deprotection, reaction products other than the support-bound peptide are removed by SRNF (see SRNFs in Chapter 8). There are many commercial firms that will supply custom-made polypeptides at a price depending on the degree of purity required.

### SOLID PHASE DEOXYRIBONUCLEOTIDE SYNTHESIS

The need for oligodeoxyribonucleotides mainly as primers for the preparation of deoxyribonucleic acids (DNA) and for DNA sequencing has resulted in considerable developments in oligo-deoxyribonucleotide synthesis. The solid phase procedure is the method commonly used. Automated synthesisers are commercially available, but with the increase in the number of firms that will provide custom-made oligo-deoxyribonucleotides, it is often not economical to purchase a synthesiser to make one's own oligo-deoxyribonucleotides. Unlike in polypeptide synthesis where there are some twenty different residues to "string" together, in DNA synthesis there are only four deoxyribonucleotides. Consequently there is usually little difficulty in synthesising 100-mers in quantities from 10 µg to 10 milligrams of material. The deprotected deoxyribonucleic acid which is separated from the solid support is purified on an anion exchange column followed by reverse phase HPLC using C8 to C18 columns for desalting. As for the polypeptides, the cost of DNA will depend on the purification level required.

### POLYMER SUPPORTED REACTANTS

These have become of increasing importance in synthesis, and a broad classification of polymer supported reactants is as follows: Polymer bound bases (e.g. dimethylaminopyridine, morpholine, piperidine); Polymer supported catalysts (e.g. Grubbs catalyst for metathesis reactions, palladium for hydrogenation reactions, tributylmethylammonium chloride for phase transfer reactions); Polymer supported condensation reagents (e.g. DEAD (diethyl azodicarboxylate) for Mitsunobu reactions, DEC [1-(3-dimethylamino-propyl)-3-ethyl-carbodiimide hydrochloride] {or EDCI [1-ethyl-3-(3-di-methylaminopropyl) carbodiimide HCl]} for peptide synthesis, HOBt (1-hydroxybenzotriazole) for peptide synthesis; Polymer supported oxidising agents (e.g. osmium tetroxide, perruthenate, pyridinium chlorochromate); Polymer supported reducing agents (e.g. borohydride, tributyltin fluoride); Polymer-supported phosphines (for miscellaneous applications depending on the structure) and so on. Commercially available polymer supported catalysts and reagents can be purchased from a number of companies including Sigma-Aldrich (see Chem Files, Vol 5, issue 11 as well as the main catalog), Biotage, Alfa-Aesar, TCI.

### SCAVENGER RESINS

The use of resins to clean up reactions has gained much importance over the years. The type of commercially available scavenger resins are electrophilic scavenger resins (e.g. benzaldehyde derivatised resins to scavenge amines; isocyanate resins to scavenge amines, anilines and hydrazines; tosyl chloride resins to scavenge nucleophiles) and nucleophilic scavenger resins (e.g. diethylenetriamine resins to scavenge acids, acid chlorides, anhydrides; sulfonyl amide resins to scavenge acids, acid chlorides, aldehydes, isocyanates and chloroformates) [e.g. see Bhattacharyya *Comb Chem High Throughput Screening* 3 65 2000, Bhattacharyya *Curr Opin Drug Disc Dev* 7 752 2004].

Specific scavengers that use purified silica support (e.g. LpDNPH cartridges which contain a high purity silica adsorbent coated with 2,4-dinitrophenylhydrazine to remove carbonyl compounds), charcoal in various forms and Florisil etc. for particular purposes are available commercially (see Biotage: the *Synthesis and Purification Catalogue* and the *Analytical Sample Preparation Catalogue* and <[www.biotage.com](http://www.biotage.com)> which contain details of such resins, and Supelco: the *Chromatographic Products for Analysis & Purification Catalogue* <[www.sigmaaldrich.com/supelco/analytical](http://www.sigmaaldrich.com/supelco/analytical)> for various supports). General types of scavenger resins are also available from a number of companies including Sigma-Aldrich and Biotage.

### COMBINATORIAL CHEMISTRY

The major impetus for the development of solid phase synthesis centers on applications in combinatorial chemistry. The notion that new drug leads and catalysts can be discovered in a high throughput fashion has been demonstrated many times over as is evidenced from the number of publications that have arisen (see references at the end of this chapter). A number of approaches to combinatorial chemistry exist. These include the split-mix method, serial techniques and parallel methods to generate libraries of compounds. The advances in combinatorial chemistry are also accompanied by sophisticated methods in deconvolution and identification of compounds from libraries. In a number of cases, innovative hardware and software have been developed for these purposes.

Depending on the size of the combinatorial library to be generated as well as the scale of the reactions to be carried out, a wide range of specialised glassware and equipment is commercially available. For example, in order to carry out parallel combinatorial synthesis, reaction stations equipped with temperature and stirring control are available from a number of sources (e.g. [www.fisher.co.uk](http://www.fisher.co.uk); [www.radleys.com](http://www.radleys.com)). These reaction stations are readily adapted, using appropriate modules, for conditions under reflux or under inert atmosphere. For automated synthesis of large libraries of compounds, reactions can be carried out using reaction blocks on microtiter plates.

## ALTERNATIVE SOLVENTS

The bulk of a reaction medium usually comprises the solvent for reaction. It has been estimated that generally solvents account for 80–90% of the mass utilisation in a typical batch reaction. In view of this, one of the greatest challenges to the ‘greening’ of chemistry is to consider the ‘greening’ of solvents. In this regard, issues such as environmental toxicities, waste minimisation, recyclability or the use of renewable solvents are considered in the greening of solvents. A number of publications have provided a framework for the environmental assessment of solvents [Capello et al. *Green Chem* **8** 927 2007, Alfonsi et al. *Green Chem* **10** 31 2008]. Some approaches to the replacement of conventional solvents include the use of supercritical fluids, e.g. carbon dioxide, ionic liquids, perfluorohydrocarbons. The use of water as an environmentally friendly reaction media has also been the focus of study by several groups [Dallinger & Kappe *Chem Rev* **107** 2563 2007, Hailes *Org Proc Res Dev* **11** 114 2007, Li & Chen *Chem Soc Rev* **35** 68 2006, Leadbeater *J Chem Soc, Chem Commun* 2881 2005] as is the study of solvent-less reactions [Polshettiwar & Varma *Pure Appl Chem* **80** 777 2008].

### SUPERCritical FLUIDS

Supercritical fluids (SCF) are defined as substances that are above their critical temperatures and pressures such that at this critical point, this is the highest temperature and pressure at which the substance can exist as a vapour and liquid in equilibrium. The commonly used SCF in chemistry is that of supercritical carbon dioxide (scCO<sub>2</sub>), which has a critical point at 31.1°C and a pressure of 73.8 Bar. The density of CO<sub>2</sub> at this critical point is approximately 0.4g/mL. CO<sub>2</sub> is potentially an ideal green solvent—it is non-toxic to the environment and at atmospheric pressure, CO<sub>2</sub> is a gas and hence does not require any waste treatment. It has a high gas-dissolving ability, a low solvation ability, good mass transfer properties and high diffusion rates. CO<sub>2</sub> is inexpensive, non-flammable, environmentally benign and can be removed completely from products. Although CO<sub>2</sub> is a greenhouse gas, CO<sub>2</sub> is naturally occurring as well as obtained as a byproduct of many processes, and the use of CO<sub>2</sub> as solvent is to use a “renewable” resource. In this respect, CO<sub>2</sub> is easily recovered either by cooling (condensing) or absorption by aqueous alkaline solutions and re-used. The disadvantages of using scCO<sub>2</sub> are the specialised and expensive equipment needed to achieve the critical conditions, the low dielectric constant and low dissolving abilities. In addition, scCO<sub>2</sub> cannot be used in reactions with strong nucleophiles. Applications with scCO<sub>2</sub> as solvent has been demonstrated, in extractions (e.g. of caffeine), dry-cleaning, polymerisation reactions, pharmaceutical processing. The use of scCO<sub>2</sub> in organic synthesis has also been demonstrated in a number of reactions including hydrogenation reactions, Diels-Alder chemistry, free radical reactions, Friedel Crafts reactions and so on [S. Hadlington *Chem Br* **39** 21 2003, Liu & Xiao, *J Mol Cat A: Chem* **270** 1 2007]. The solvent properties of scCO<sub>2</sub> can be tuned by adjusting the temperature and pressure. Liquids and solid solutes are surprisingly soluble in scCO<sub>2</sub> though it is generally true that the solubility of non-polar substances is higher than the polar ones because CO<sub>2</sub> is a non-polar molecule. The solubility can be tuned by changing the bulk densities of CO<sub>2</sub> or by adding a co-solvent or modifying the solute. Other fluids that have been used in this way are ammonia and SO<sub>2</sub>, although they do not have the many advantages of CO<sub>2</sub>, especially those pertaining to its low reactivity under “normal” reaction conditions.

### IONIC LIQUIDS

Ionic liquids (ILs) are organic or inorganic salts that are liquids at room or reaction temperatures. Although ionic liquids are themselves not new discoveries (e.g. the ionic liquid [EtNH<sub>3</sub>][NO<sub>3</sub>] was described in 1914), the use of ionic liquids in synthesis is relatively recent. In particular, the potential applications of ionic liquids as solvents in synthesis and in catalysis have been achieved. They have high mechanical, thermal and electrochemical stability. Their physical properties make them unique solvents for synthesis, and are ‘green’ alternatives to volatile organic solvents, though there has been some debate on whether ionic liquids can be considered as ‘green’ solvents. [R. Sanghi & M.M. Srivasatava, *Green Chemistry: Environment Friendly Alternatives*, Alpha Science International, 2003, ISBN 9781842651735; C. Alfonso, A.M. Crespo & P.S.G. Joao, *Green Separation Processes: Fundamentals and Applications*, VCH-Wiley Publishing, 2005, ISBN 978357309856; J. Clark & J. Hardy, *Green Chemistry in Undergraduate Practical Classes (with CD-ROM)*, Royal Society of Chemistry, 2006, ISBN 978054042784]. The impact of ILs on health and environment is not known. In addition to this, many ILs are difficult to prepare and can be expensive if bulk, high purity ILs are needed. Ionic liquids are good solvents for both organic and inorganic substances and hence can be used to bring reagents into the same phase for reaction. Ionic liquids are also immiscible with a number of organic solvents and thus provide a non-aqueous, polar alternative for two-phase extraction systems. As ionic liquids are non-flammable and non-volatile (with hardly any measurable vapour pressure), they can be used in high vacuum systems without the possibility of loss or contaminants. In addition, this also facilitates the isolation of products as products can be distilled from the ionic liquid or alternatively extracted with an organic solvent that is immiscible with the ionic liquid. Although ionic liquids are frequently composed of poorly coordinating ions,

they are highly polar which are important characteristics in the activation of catalysts. For example the 1-butyl-3-methylimidazolium salt with  $\text{AlCl}_3$  is a very good catalyst for the Friedel-Crafts reaction, and varying the ratio of salt to  $\text{AlCl}_3$  produces a more or less acidic medium. ILs are used successfully in organic reactions which are mediated by microwaves. ILs have found uses in phase-separation processes, in electrochemical processes, as heat storage media, lubricants and additives.

A large variety of ILs of high purity with water content below 100ppm and halide content below 10ppm are now available commercially. These include numerous *N*-alkylpyridinium, *N,N'*-dialkylimidazolium, alkylammonium and alkylphosphonium salts, covalent hydrophobic ILs (e.g. 1,2-dimethyl-3-propylimidazolium bis(trifluoromethylsulfonyl)imide, task-specific ILs. Bronsted acidic ILs (e.g. 3-[triphenylphosphonio]propane-1-sulfonic acid tosylate), nitrile-functionalised ILs, perfluorinated ILs (e.g. tetrabutylammonium nonafluorobutane sulfonate, BASF Basonics (e.g. 1,2,3-trimethylimidazolium methyl sulfate), and TOMATS for heavy metal extraction (e.g. methyltrioctylammonium thiosalicylate)

A number of reactions have been carried out in ionic liquids; for examples see Dell'Anna et al. *J Chem Soc, Chem Commun* 434 2002, Nara et al. *Tetrahedron Lett* 43 1127 2002, Semeril et al. *J Chem Soc, Chem Commun* 146 2002, Buijsman et al. *Org Lett* 3 3785 2001. These include Diels-Alder reactions, Friedel-Crafts (above), transition-metal mediated catalysis, e.g. Heck and Suzuki coupling reactions, and olefin metathesis reactions. An example of ionic liquid acceleration of reactions carried out on solid phase is given by Revell & Ganesan [*Org Lett* 4 3071 2002]. [See also Fluka, Riedel-deHaën, Catalogue 2007–2008, and <www.sigmaaldrich.com/ionicliquids>, with or without Basonics™ and Cytec (Cytos<sup>R</sup>); Chem Files (Sigma-Aldrich) Vol 6, No 9 2006, Ghassemi et al. Application Note 48 in Synthesis and Purification Catalog 2007, <www.Biotage.com>.]

### FLUOROUS CHEMISTRY

Fluorous chemistry is defined as chemistry that is related to highly fluorinated  $\text{sp}^3$ -hybridised carbon containing compounds. The seminal work by Horvath and Rabai (*Science* 266 72 1994) described the applications of fluorous biphasic catalysis and applied it to Rh catalysed hydroformylation reactions. To date the scope of fluorous chemistry has expanded to include applications in organic synthesis and material science, separation, extraction and chromatography. For example, fluorous phase labels can be attached to substrates such that the subsequent fluorinated products can be extracted into the fluorous phase. This has applications in liquid-liquid extractions (typical work-up procedures) where a three-phase extraction is now possible (organic, fluorous and aqueous phases). As organic and inorganic compounds have little or no tendency to dissolve in highly fluorinated solvents and compounds, phase labelling a compound as fluorous will enable successful extraction into the fluorous phase. However, in order to carry out homogenous reactions with these fluorinated compounds, organic solvents with a good dissolving power for fluorous compounds or miscible organic and fluorous solvents can be used. Alternatively organic solvents with a few fluorine atoms, e.g. trifluoroethanol, benzotrifluoride ("hybrid solvents") will dissolve both organic and fluorous compounds. A number of synthetic applications utilising fluorous chemistry have been reported in the literature. [For examples, see Schneider & Bannwarth *Helv Chim Acta* 84 735 2001, Galante et al. *Tetrahedron Lett* 42 5425 2001, Studer & Curran *Tetrahedron* 53 6681 1997, Studer et al. *J Org Chem* 62 2917 1997, Crich & Neelamkavil *Tetrahedron* 58 3865 2002.] For some fluorous compounds for synthesis and separation, see Chem Files from Sigma-Aldrich, Supplement II, 2008. One potential limitation of fluorous solvents is their cost.

### BIOMASS DERIVED ORGANIC SOLVENTS

Biomass derived organic solvents are solvents that are obtained from renewable resources, which is in contrast to the traditional petrochemical based solvents that are limited in resource. These solvents are gaining traction as green solvents, though the scope of the reaction transformations and other solvent replacement applications have not been fully studied. Examples of such biomass derived solvents are 2-methyltetrahydrofuran [96-47-9], glycerol [56-81-5], glycerol ethers, ethyl lactate and cyclopentylmethyl ethers. 2-Methyltetrahydrofuran is an aprotic ether which has polarity and Lewis base strengths between that for diethyl ether and THF [Aycock *Org Proc Res Dev* 11 156 2007; Ripin & Vetelino *Syn Lett* 2353 2003]. It had been used as a replacement solvent for THF [109-99-9] in organometallic reactions (e.g. Grignard, lithiation, hydride reduction, coupling reactions) with the advantage that MeTHF is only partially water miscible and hence can be conveniently used to recover reaction products [Comanita & Aul *Chimica Oggi* 25 26 2007; Pace et al. *Tetrahedron* 67 2670 2011]. Cyclopentylmethyl ether has been used as an alternative to ethereal solvents such as THF [109-99-0], 1,4-dioxane [123-91-1], 1,2-dimethoxyethane [110-71-4] [Watanabe et al. *Org Proc Res Dev* 11 251 2007]. The formation of peroxides is suppressed in cyclopentylmethyl ether [5614-37-9] and this solvent has a narrow explosion range and can be easily dried. The use of cyclopentylmethyl ether as solvent has also been demonstrated in a number of reactions including organometallic reactions, oxidation and reduction reactions

and reactions with transition metal catalysts. Interestingly, a promising new sustainable solvent is glycerol. Glycerol is an organic waste product from biodiesel production and hence is cheap, readily available and is renewable. Glycerol is a non-toxic polar solvent that is able to dissolve inorganic salts, acids, bases, transition metal complexes as well as organic compounds that are poorly miscible in water [Diaz-Alvarez *Chem Commun* **47** 6208 2011; Gu & Jerome *Green Chem* **12** 1127 2010]. The disadvantages of using glycerol as solvent are the high viscosity as well as the possible reactivity of the hydroxyl groups in chemical transformations. Nevertheless, the use of glycerol as solvent in a variety of organic transformations has been demonstrated. In some cases, glycerol as solvent has a beneficial effect on reaction rates and selectivity. Due to glycerol's immiscibility with non-polar solvents, the use of glycerol facilitates isolation of reaction products and has been used in separation processes.

## BIBLIOGRAPHY

### GREEN CHEMISTRY

P.T. Anastas and J.C. Warner, *Green Chemistry: Theory and Practice*, Oxford University Press, USA 2005. ISBN10: 0198506988, ISBN-13: 978-0198506980.

R.A. Sheldon, I. Arends and U. Hanefeld, *Green Chemistry and Catalysis*, Wiley-VCH, 2007. ISBN 10: 352730715x, ISBN-13: 978-3527307159.

P. Tundo, A. Perosa and F. Zecchini, *Methods and Reagents for Green Chemistry: An Introduction*. Wiley-Interscience, 2007. ISBN-10: 0471754005, ISBN-13: 978-0471754008.

### ATOM ECONOMY

H.M.L. Davies and D. Morton, C-C bond formation by C-H bond activation in *Science of Synthesis, Stereoselective Synthesis* (J.G. De Vries, G.A. Molander and A.P. Evans Eds) **3** 513-566 2011.

J. Du Bois, Rhodium catalyzed C-H amination: an enabling method for chemical synthesis, *Organic Process Research and Development* **15** 758-762 2011.

G. Dyker, Transition metal catalyzed coupling reactions under C-H activation, *Angew Chem. Int Ed* **38** 1698-1712 1999.

C. Liana Allen and J.M.J. Williams, Metal catalyzed approaches to amide bond formation, *Chem Soc Rev* **40** 3405-3415 2011.

K. Goldberg and A.S. Goldberg, in Activation and Functionalisation of C-H bonds, *ACS Symposium Series*, Oxford University Press, 2004.

S. Murai, in Activation of Unreactive bonds and Organic Synthesis, *Topics in Organometallic Chemistry*, Springer-Verlag, Berlin Heidelberg, 1999, Vol 3.

T. Newhouse and P.S. Baran, If C-H bonds could talk: selective C-H bond oxidation, *Angew Chem, Int Ed* **50** 3362-3374 2011.

V. Ritleng, C. Sirlin, and M. Pfeffer, Ru-, Rh-, and Pd catalyzed C-C bond formation involving C-H activation and addition on unsaturated substrates: reactions and mechanistic aspects, *Chem Rev* **102** 1731-1769 2002.

D.R. Stuart and K. Fagnou, The catalytic cross coupling of unactivated arenes, *Science* **316** 1172-1175 2007.

D. Zhao, J. You and C. Hu, Recent Progress in coupling of two heteroarenes, *Chemistry – A Eur J* **17** 5466-5492 2011.

## ORGANOCATALYSIS

A. Berkessel and H. Gröger, *Asymmetric Organocatalysis: From Biomimetic Concepts to Applications in Asymmetric Synthesis*, Wiley-VCH 2005. ISBN-10: 3-527-30517-3, 13: 9783527305179.

S. Bertelsen and K.A. Jorgensen, Organocatalysis-after the gold rush, *Chem Soc Rev* **38** 2178-2189 2009.

P.I. Dalko (Ed.), *Enantioselective Organocatalysis: Reactions and Experimental Procedures*, Wiley-VCH, 2007. ISBN-10: 3527315225, ISBN-13: 978-3527315222.

R.M. deFigueiredo and M. Christmann, Organocatalytic synthesis of drugs and bioactive natural products, *Eur J Org Chem* 2575-2600 2007.

D. Enders, C. Grondal and M.R. Huttl, Asymmetric organocatalytic domino reactions, *Angew Chem, Int Ed* **46** 1570 2007.

E.N. Jacobsen and D.W.C. MacMillan, Organocatalysis, *Proc Natl Acad Sci* **107** 20618-20619 2010.

D. Kampen, C.M. Reisinger and B. List, Chiral Bronsted acids for asymmetric organocatalysis, *Topics in Current Chemistry* (I. Ojima Ed) **291** 395-456 2010.

G. Lelais and D.W. MacMillan, Modern strategies in organic catalysis: the advent and development of iminium activation, *Aldrichimica Acta* **39** 79-87 2006.

B. List, The ying and yang of organic asymmetric aminocatalysis, *Chem Commun* 819 2006.

H. Pellissier, Asymmetric organocatalysis, *Tetrahedron* **63** 9267-9331 2007.

J. Seayad and B. List, Asymmetric organocatalysis, *Org Biomol Chem* **3** 719-724 2005.

O. Sereida, S. Tabassum, and R. Wilhelm, Lewis acid organocatalysts, *Topics in Current Chemistry* (I. Ojima Ed) **291** 349-393 2010.

A. Ting, J.M. Goss, N.T. McDougal and S.E. Schaus, Bronsted base catalysts, *Topics in Current Chemistry* (I. Ojima Ed) **291** 145-200 2010.

C. Zhong and X. Shi, When organocatalysts meets transition metal catalysis. *Eur J Org Chem* 2999 2010.

## MICROWAVE TECHNOLOGIES

J.A. Dahl, B.L.S. Maddux and J.E. Hutchison, Towards greener nanosynthesis, *Chem Rev* **107** 2228-2269 2007.

A. de la Hoz, A. Diaz-Ortiz and A. Moreno, Microwaves in organic synthesis: Thermal and non-thermal microwave effects, *Chem Soc Rev* **34** 164-178 2005.

B.L. Hayes, *Microwave Synthesis: Chemistry at the Speed of Light*, CEM publishing, Matthews NC, 2002. ISBN-10: 092222901, ISBN-13: 978-0972222907.

O.C. Kappe, Synthetic methods Controlled microwave heating in modern organic synthesis, *Angew Chem, Int Ed* **43** 6250-6284 2004.

O.C. Kappe, The use of microwave irradiation in organic synthesis. From Laboratory Curiosity to Standard Practice in Twenty Years, *Chimia* **60** 308-312 2006.

P. Lindström, J.P and Tierney (Eds.), *Microwave-assisted Organic Synthesis*, Blackwell, Oxford or CRC Press, 2005. ISBN-10: 0849 323711, ISBN-13: 978-0849323713.

A. Loupy (Ed.) *Microwaves in Organic Synthesis* (2 volume set), Wiley-VCH, 2006. ISBN-10: 3527314520, ISBN-13: 978-3527314522.

V. Polshettiwar and R.S. Varma, Microwave-assisted organic synthesis and transformations using benign reaction media, *Acc Chem Res* **41** 629-639 2008.

## SOLID PHASE SYNTHESIS

### General

J.A. Ellman, Design, synthesis, and evaluation of small-molecule libraries, *Acc Chem. Res* **29** 132-143 1996.

J.S. Fruchtel and G. Jung, Organic chemistry on solid supports, *Angew Chem, Int Ed* **35** 17-42 1996.

A. Ganesan, Solid-phase synthesis in the Twenty-first century, *Mini-reviews in Medicinal Chemistry* **6** 3-10 2006.

F. Guillier, D. Orain and M. Bradley, Linkers and cleavage strategies in solid-phase organic synthesis and combinatorial chemistry, *Chem Rev* **100** 2091-2157 2000.

P. Hodge, Polymer-supported organic reactions: What takes place in the beads? *Chem Soc Rev* **26** 417-424 1997.

J.W. Labadie, Polymeric supports for solid phase synthesis, *Curr Opin Chem Biol* **2** 346-352 1998.

P. Seneci, *Solid Phase and Combinatorial Technologies*, J. Wiley & Sons, New York, 2000, ISBN 0471331953.

A.R. Vaino and K.D. Janda, Solid-phase organic synthesis: A critical understanding of the resin, *J Comb Chem* **2** 579-596 2000.

### Applications to SPOS, SPPS

P. Blaney, R. Grigg and V. Sridharan, Traceless solid-phase organic synthesis, *Chem Rev* **102** 2607-2624 2002.

B. Clapham, T.S. Reger and K.D. Janda, Polymer-supported catalysis in synthetic organic chemistry, *Tetrahedron* **57** 4637-4662 2001.

D.H. Drewry, D.M. Coe and S. Poon, Solid supported reagents in organic synthesis, *Med Res Rev* **19** 97-148 1999.

S.J. Shuttleworth, S.M. Allin and P.K. Sharma, Functionalised polymers –recent developments and new applications in synthetic organic chemistry, *Synthesis* 1217-1239 1997.

S.J. Shuttleworth, S.M. Allin, R.D. Wilson and D. Nasturica, Functionalised polymers in organic chemistry; Part 2, *Synthesis* **8** 1035-1074 2000.

### Polymer Supported Reagents And Scavenger Resins

R.J. Booth and J.C. Hodges, Solid-supported reagent strategies for rapid purification of combinatorial synthesis products, *Acc Chem Res* **32** 18-26 1999.

G.J. Gabriel, Scavenger Resins in solution-phase combi-chem, *Methods Enzymol* **369** (Combinatorial Chemistry Part B) 391-412 2003.

S.V. Ley, I.R. Baxendale, G. Brusotti, M. Caldarelli, A. Massi and M. Nesi, Solid-supported reagents for multi-step organic synthesis: preparation and application, *Il Farmaco* **57** 321-330 2002.

H. Salimi, A. Rahimi, and A. Pourjavadi, Applications of polymeric reagents in organic synthesis, *Monatsh Chem* **138**, 363-379 2007.

A. Solinas and M. Taddei, Solid-supported reagents and catch-and-release techniques in organic synthesis, *Synthesis* 2409-2453 2007.

P.H. Toy and M. Shi, Polymer-supported reagents and catalysts: increasingly important tools for organic synthesis, *Tetrahedron* **61**, 12025 2005.

### Combinatorial Chemistry

F. Balkenhohl, C. Vondembusschehunnefeld, A. Lansky and C. Zechel, Combinatorial synthesis of small organic molecules, *Angew Chem, Int Ed* **35** 2289-2337 1996.

A. Ganesan, Recent developments in combinatorial organic synthesis, *Drug Discovery Today* **7** 47-55 2002.

H. Wennemers, Combinatorial chemistry: a tool for the discovery of new catalysts, *Comb Chem High Throughput Screening* **4** 273-285 2001.

### SUPERCRITICAL CARBON DIOXIDE

E.J. Beckman, Supercritical and near critical CO<sub>2</sub> in green chemistry synthesis and processing. *J Supracrit Fluids* **28** 121-191 2004.

P. Licence, J. Ke, M. Sokolova, S.K. Ross and M. Poliakoff, Chemical reactions in supercritical carbon dioxide: from laboratory to commercial plant, *Green Chem* **5** 99-104 2003.

W. Leitner, Carbon dioxide as an environmentally benign reaction medium for chemical synthesis, *Appl Organomet Chem* **14** 809-814 2000.

W. Leitner, Supercritical carbon dioxide as a green reaction medium for catalysis, *Acc Chem Res* **35** 746-756 2002.

R.S. Oakes, A.A. Clifford and C.M. Rayner, The use of supercritical fluids in synthetic organic chemistry, *J Chem Soc, Perkins Trans 1* 917-941 2001.

C.M. Rayner, The potential of carbon dioxide in synthetic organic chemistry, *Org Proc Res Dev* **11** 121-132 2007.

### IONIC LIQUIDS

C.M. Gordon, New developments in catalysis using ionic liquids, *Appl Catal A: General* **222** 101-117 2001.

W. Miao and T.H. Chan, Ionic liquid-supported synthesis: A novel liquid-phase strategy for organic synthesis, *Acc Chem Res* **39** 897-908 2006.

R. Sheldon, Catalytic reactions in ionic liquids, *J Chem Soc, Chem Commun* 2399-2407 2001.

P. Wasserscheid and T. Welton, *Ionic Liquids in Synthesis* (2 volume set), Wiley VCH, 2007. ISBN-10: 3527312390, ISBN-13: 978-3527312399.

T. Welton, Room temperature ionic liquids. solvents for synthesis and catalysis, *Chem Rev* **99** 2071-2083 1999.

## FLUOROUS CHEMISTRY

D.P. Curran, Combinatorial organic synthesis and phase separation: Back to the future, *Chemtracts: Org Chem* **9** 75-87 1996.

D.P. Curran, Strategy-level separations in organic synthesis: from planning to practice, *Angew Chem, Int Ed* **37** 1174-1196 1998.

D.L. Flynn, Phase-trafficking reagents and phase-switching strategies for parallel synthesis, *Med Res Rev* **19** 408-431 1999.

J.A. Gladysz, D.P. Curran and I.T. Horvath (Eds.), *Handbook of Fluorous Chemistry*, Wiley VCH, 2004. ISBN-10: 352730617x, ISBN-13: 978-3527306176.

A. Studer, S. Hadida, R. Ferritto, S.Y. Kim, P. Jeyer, P. Wipf and D.P. Curran, Fluorous synthesis — A fluoros-phase strategy for improving separation efficiency in organic synthesis, *Science* **275** 823-826 1997.

J.A. Gladysz and D.P. Curran, *Tetrahedron* **58** Issue 20, pp 3823-4132 2002.

W. Zhang and D.P. Curran, Synthetic applications of fluoros solid-phase extraction (F-SPE), *Tetrahedron* **62** 11837-11865 2006.

# CHAPTER 4

## PURIFICATION OF ORGANIC CHEMICALS

### INTRODUCTION

The general principles, techniques and methods of purification in Chapters 1 and 2 are applicable to this chapter. Most organic liquids and a number of solids can readily be purified by fractional distillation, usually at atmospheric pressure. Sometimes, particularly with high boiling or sensitive liquids, or when in doubt about stability, distillation or fractionation under as low a pressure as possible should be carried out. To save space, the present chapter omits many substances for which the published purification methods involve simple distillation. Where boiling points are given, purification by distillation is another means of removing impurities. Literature references, and in particular Beilstein references, are included for most entries which refer the reader directly or indirectly to the original sources. Substances are listed alphabetically in each section, usually with some criteria of purity, giving brief details of how they can be purified. Also noted are the molecular weights (to the first decimal place), melting points and/or boiling points together with the respective densities and refractive indexes for liquids, and optical rotations for chiral compounds. All temperatures are in centigrade unless stated otherwise. When temperatures and/or wavelengths are not given for the last three named properties, then they should be assumed to be 20°C and the average of the wavelengths of the sodium D lines respectively; and most densities are relative to water at 4°C.

**Ionisation constants** of ionisable compounds are given as **pK** values (published in the literature) and refer to the **pK<sub>a</sub>** values at room temperature (~15°C to 25°C). Values at other temperatures are given as superscripts, e.g. **pK<sup>30</sup>** for 30°C. Estimated values are entered as **pK<sub>Est</sub>** (see Chapter 1, pp 34–36).

The present chapter includes commercially available organic chemicals. Most of the inorganic, metal-organic, organo- bismuth, boron, phosphorus, selenium, silicon and alkali metal compounds and metal ion salts of organic acids are in Chapter 5. Naturally occurring commercially available organic compounds for use in biochemistry, molecular biology and biology are in Chapter 6. Commercially available polymer supported reagents are indicated with § under the appropriate reagent.

**Rapid purification procedures** are included for commonly used solvents and reagents which make them suitable for general use in synthetic chemistry.

**Abbreviations** of titles of periodicals are defined as in the Chemical Abstracts Service Source Index (CASSI) except that punctuation is deleted. Other abbreviations are self evident.

**As a good general rule, all low boiling (<100°) organic liquids should be treated as highly flammable and toxic (because they can be inhaled in large quantities) and the necessary precautions should be taken (see Safety precautions associated with the purification of laboratory chemicals in Chapter 1, p 4).**

**Benzene** has been used as a solvent successfully and extensively in the past for reactions and purification by chromatography and crystallisation but is now considered a **very dangerous substance**. It should be used with extreme care. We emphasise that an alternative solvent to benzene (e.g. toluene, toluene-petroleum ether, or a petroleum ether to name a few) should be used first. However, if no other solvent system can be found, then all operations involving benzene (and storage) have to be performed in an efficiently running fumehood, and precautions must be taken to avoid inhalation and contact with skin and eyes. An asterisk has been inserted in the text, e.g. \*C<sub>6</sub>H<sub>6</sub> or \*benzene, to remind the user that special precautions should be adopted.

This chapter consists of five sections *viz*: **Aliphatic Compounds, Alicyclic Compounds, Aromatic Compounds, Heterocyclic Compounds** and **Miscellaneous As, B, P, Si, S, Se, and Te Compounds**.

## ALIPHATIC COMPOUNDS

**Acetal (acetaldehyde diethylacetal)** [105-57-7] **M 118.2, b 103.7-104°, d<sub>4</sub><sup>20</sup> 0.831, n<sub>D</sub><sup>25</sup> 1.38054, n<sub>D</sub><sup>25</sup> 1.3682.** Dry acetal over Na to remove alcohols and H<sub>2</sub>O, and to polymerise aldehydes, then fractionally distil. Or, treat it with alkaline H<sub>2</sub>O<sub>2</sub> at 40-45° to remove aldehydes, then saturate with NaCl, separate, dry with K<sub>2</sub>CO<sub>3</sub> and distil it from Na [Vogel *J Chem Soc* 616 1948]. [Beilstein 1 IV 3103.]

**Acetaldehyde** [75-07-0] **M 44.1, b 20.2°, d<sub>4</sub><sup>20</sup> 0.788, n<sub>D</sub><sup>25</sup> 1.33113, pK<sup>25</sup> 13.57 (hydrate).** Acetaldehyde is usually purified by fractional distillation in a glass helices-packed column under dry N<sub>2</sub>, discarding the first portion of distillate. Or, it is shaken for 30 minutes with NaHCO<sub>3</sub>, dried with CaSO<sub>4</sub> and fractionally distilled at 760mm through a 70cm Vigreux column. The middle fraction is collected and further purified by standing for 2 hours at 0° with a small amount of hydroquinone (free radical inhibitor), followed by distillation [Longfield & Walters *J Am Chem Soc* 77 810 1955]. [Beilstein 1 IV 3094.]

**Acetaldehyde dimethyl acetal (1,1-dimethoxyethane)** [534-15-6] **M 90.1, b 63-65°, d<sub>4</sub><sup>20</sup> 0.852, n<sub>D</sub><sup>25</sup> 1.36678.** Distil the dimethyl acetal through a fractionating column and fraction boiling at 63.8°/751mm is collected. It forms an azeotrope with MeOH. *Alternatively*, purify it as for *acetal* above. It has been purified by GLC. [Beilstein 1 IV 3103.]

**Acetamide** [60-35-5] **M 59.1, m 81°, pK<sub>1</sub><sup>25</sup> -1.4, pK<sub>2</sub><sup>25</sup> +0.37.** Acetamide is crystallised by dissolving in hot MeOH (0.8ml/g), diluting with Et<sub>2</sub>O and allowing to stand [Wagner *J Chem Edu* 7 1135 1930]. Alternate crystallisation solvents are acetone, \*benzene, chloroform, dioxane, methyl acetate or \*benzene/ethyl acetate mixtures (3:1 and 1:1). It has also been recrystallised from hot water after treating with HCl-washed activated charcoal (which had been repeatedly washed with water until free from chloride ions), then crystallised again from hot 50% aqueous EtOH and finally twice from hot 95% EtOH [Christoffers & Kegeles *J Am Chem Soc* 85 2562 1963]. Finally it is dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. Acetamide is also purified by distillation (**b** 221-223°) or by sublimation *in vacuo*. It has also been purified by two recrystallisations from cyclohexane containing 5% (v/v) of \*benzene. Needle-like crystals separate and are filtered, washed with a small volume of distilled H<sub>2</sub>O and dried with a flow of dry N<sub>2</sub>. [Slebocka-Tilk et al. *J Am Chem Soc* 109 4620 1987, Beilstein 2 H 175, 2 I 80, 2 II 177, 2 III 384, 2 IV 399.]

**Acetamidine hydrochloride** [124-42-5] **M 94.5, m 164-166°, 165-170°(dec), 174°, pK<sup>25</sup> 12.40.** The hydrochloride can be recrystallised from small volumes of EtOH. *Alternatively*, it is dissolved in EtOH, filtered, Et<sub>2</sub>O is added; filter the crystalline salt off under N<sub>2</sub> and dry it in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. The salt is deliquescent and should be stored in a tightly stoppered container. Its solubility in H<sub>2</sub>O is 10% at room temperature and it is soluble in Me<sub>2</sub>CO. The *free base* reacts strongly alkaline in H<sub>2</sub>O. It has λ<sub>max</sub> 224nm (ε 4000) in H<sub>2</sub>O. The *picrate* has **m** 252° (sintering at ~245°). [Dox *Org Synth Coll Vol I* 5 1941, Davies & Parsons *Chem Ind (London)* 628 1958, Barnes et al. *J Am Chem Soc* 62 1286 1940 give **m** 177-178°, Beilstein 2 H 185, 2 I 85, 2 II 183, 2 III 416, 2 IV 428.]

**N-(2-Acetamido)-2-aminoethanesulfonic acid (ACES)** [7365-82-4] **M 182.2, m > 220°(dec), pK<sub>Est</sub> ~1.5, pK<sub>2</sub> 6.9.** Recrystallise ACES from hot aqueous EtOH. [Perrin & Dempsey *Buffers for pH and Metal Ion Control* Chapman & Hall, London 1974, Beilstein 4 III 1707.]

**N-(2-Acetamido)iminodiacetic acid (ADA)** [26239-55-4] **M 190.2, m 219°(dec), pK<sub>1</sub> ~2.3, pK<sub>2</sub> 6.6.** Dissolve ADA in water, add one equivalent of NaOH solution (to final pH of 8-9), then acidify with HCl to precipitate the free acid. This is filtered off, washed with water and dried *in vacuo*. [Beilstein 4 IV 2441.]

**Acetamidomethanol** [625-51-4] **M 89.1, m 47-50°, 54-56°, 55°.** Recrystallise it from freshly distilled Me<sub>2</sub>CO, wash the crystals with dry Et<sub>2</sub>O and dry them in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. R<sub>F</sub> 0.4 on paper chromatography with CHCl<sub>3</sub>/EtOH (2:8) as solvent and developed with ammoniacal AgNO<sub>3</sub>. It also crystallises in needles from EtOAc containing a few drops of Me<sub>2</sub>CO. It is *hygroscopic* and should be stored under dry conditions. [Bachmann et al. *J Am Chem Soc* 73 2775 1951, Walter et al. *Chem Ber* 99 3204 1966, Einhorn &

Ladisch *Justus Liebigs Ann Chem* **343** 265 1905, *Beilstein* **2** IV 405.]

**Acetic acid (glacial)** [64-19-7] **M 60.1, m 16.6°, b 118°, d<sub>4</sub><sup>20</sup> 1.049, n<sub>D</sub><sup>25</sup> 1.37171, n<sub>D</sub><sup>25</sup> 1.36995, pK<sub>25</sub><sup>25</sup> 4.76.** Usual impurities are traces of acetaldehyde and other oxidisable substances and water. (Glacial acetic acid is very *hygroscopic*. The presence of 0.1% water lowers its **m** by 0.2°.) Purify it by adding some acetic anhydride to react with water present, heat it for 1 hour to just below boiling in the presence of 2g CrO<sub>3</sub> per 100ml and then fractionally distil it [Orton & Bradfield *J Chem Soc* 960 1924, Orton & Bradfield *J Chem Soc* 983 1927]. Instead of CrO<sub>3</sub>, use 2-5% (w/w) of KMnO<sub>4</sub>, and boil under reflux for 2-6 hours. Traces of water have been removed by refluxing with tetraacetyl diborate (prepared by warming 1 part of boric acid with 5 parts (w/w) of acetic anhydride at 60°, cooling, and filtering off, followed by distillation [Eichelberger & La Mer *J Am Chem Soc* **55** 3633 1933].

Refluxing with acetic anhydride in the presence of 0.2g % of 2-naphthalenesulfonic acid as catalyst has also been used [Orton & Bradfield *J Chem Soc* 983 1927]. Other suitable drying agents include anhydrous CuSO<sub>4</sub> and chromium triacetate: P<sub>2</sub>O<sub>5</sub> converts some acetic acid to the anhydride. Azeotropic removal of water by distillation with thiophene-free benzene or with butyl acetate has been used [Birdwhistell & Griswold *J Am Chem Soc* **77** 873 1955]. An alternative purification uses fractional freezing. [*Beilstein* **2** H 96, **2** IV 94.]

**Rapid procedure:** Add 5% acetic anhydride, and 2% of CrO<sub>3</sub>. Reflux and fractionally distil.

**Acetic anhydride** [108-24-7] **M 102.1, b 138°, d<sub>4</sub><sup>20</sup> 1.082, n<sub>D</sub><sup>25</sup> 1.3904.** Adequate purification can usually be achieved by fractional distillation through an efficient column. Acetic acid can be removed by prior refluxing with CaC<sub>2</sub> or with coarse Mg filings at 80-90° for 5 days, or by distillation from a large excess of quinoline (1% AcOH in quinoline) at 75mm pressure. Acetic anhydride can also be dried by standing with Na wire for up to a week, removing the Na and distilling it under vacuum. (Na reacts vigorously with acetic anhydride at 65-70°). Dippy & Evans [*J Org Chem* **15** 451 1950] let the anhydride (500g) stand over P<sub>2</sub>O<sub>5</sub> (50g) for 3 hours, then decanted it and stood it with ignited K<sub>2</sub>CO<sub>3</sub> for a further 3 hours. The supernatant liquid was distilled and the fraction **b** 136-138° was further dried with P<sub>2</sub>O<sub>5</sub> for 12 hours, followed by shaking with ignited K<sub>2</sub>CO<sub>3</sub>, before two further distillations through a five-section Young and Thomas fractionating column. The final material distilled at 137.8-138.0°. It can also be purified by azeotropic distillation with toluene: the azeotrope boils at 100.6°. After removal of the remaining toluene, the anhydride is distilled [sample had a specific conductivity of 5×10<sup>-9</sup> ohm<sup>-1</sup>cm<sup>-1</sup>]. [*Beilstein* **2** H 96, **2** I 39, **2** II 91, **2** III 134, **2** IV 94.]

**Rapid procedure:** Shake with P<sub>2</sub>O<sub>5</sub>, separate, shake with dry K<sub>2</sub>CO<sub>3</sub> and fractionally distil.

**Acetic hydrazide** [1068-57-1] **M 74.1, m 67°, b 127°/18mm.** Acetic hydrazide crystallises as needles from EtOH. It reduces NH<sub>3</sub>/AgNO<sub>3</sub>. [*Beilstein* **2** H 191, **2** IV 435.]

**Acetoacetamide** [5977-14-0] **M 101.1, m 54-55°, 54-56°.** Recrystallise the amide from CHCl<sub>3</sub>, or Me<sub>2</sub>CO/petroleum ether. It also crystallises from pyridine with 4mols of solvent. It is slightly soluble in H<sub>2</sub>O, EtOH and AcOH but is insoluble in Et<sub>2</sub>O. The *phenylhydrazone* has **m** 128°. [Kato *Chem Pharm Bull Jpn* **15** 921 923 1967, Claisen & Meyer *Chem Ber* **35** 583 1902, *Beilstein* **3** H 659, **3** I 231, **3** III 1204, **3** IV 1545.]

**Acetone** [67-64-1] **M 58.1, b 56.2°, d<sub>4</sub><sup>20</sup> 0.791, n<sub>D</sub><sup>25</sup> 1.35880, pK<sub>1</sub><sup>25</sup> -6.1 (basic, mono-protonated), pK<sub>2</sub><sup>25</sup> 20.0 (acidic)** The commercial preparation of acetone by catalytic dehydrogenation of isopropyl alcohol gives relatively pure material. Analytical reagent quality generally contains less than 1% of organic impurities but may have up to about 1% of H<sub>2</sub>O. Dry acetone is appreciably *hygroscopic*. The main organic impurity in acetone is mesityl oxide, formed by aldol condensation. It can be dried with anhydrous CaSO<sub>4</sub>, K<sub>2</sub>CO<sub>3</sub> or type 4A Linde molecular sieves, and then distilled. Silica gel and alumina, or mildly acidic or basic desiccants cause acetone to undergo the aldol condensation, so that its water content is increased by passage through these reagents. This also occurs to some extent when P<sub>2</sub>O<sub>5</sub> or sodium amalgam is used. Anhydrous MgSO<sub>4</sub> is an inefficient drying agent, and CaCl<sub>2</sub> forms an addition compound. Drierite (anhydrous CaSO<sub>4</sub>) offers minimum acid and base catalysis for aldol formation and is the recommended drying agent for this solvent [Coetzee & Siao *Inorg Chem* **14** 2 1987, Riddick & Bunger *Organic Solvents* Wiley-Interscience, N.Y., 3rd edn, 1970]. Acetone can be shaken with Drierite (25g/L) for several hours before it is decanted and distilled from fresh Drierite (10g/L) through an efficient column, maintaining atmospheric contact through a Drierite drying tube.

The equilibrium water content is about  $10^{-2}$ M. Anhydrous  $\text{Mg}(\text{ClO}_4)_2$  **should not be used as drying agent as there is a high risk of EXPLOSION with acetone vapour.**

Organic impurities have been removed from acetone by adding 4g of  $\text{AgNO}_3$  in 30ml of water to 1L of acetone, followed by 10ml of M NaOH, shaking for 10 minutes, filtering, drying with anhydrous  $\text{CaSO}_4$  and distilling [Werner *Analyst (London)* **58** 335 1933]. Alternatively, successive small portions of  $\text{KMnO}_4$  have been added to acetone at reflux, until the violet colour persists, followed by drying and distilling. Refluxing with chromium trioxide ( $\text{CrO}_3$ ) has also been used. Methanol has been removed from acetone by azeotropic distillation (at  $35^\circ$ ) with methyl bromide, and treatment with acetyl chloride.

Small amounts of acetone can be purified as the NaI addition compound, by dissolving 100g of finely powdered NaI in 400g of boiling acetone, then cooling in ice and salt to  $-8^\circ$ . Crystals of  $\text{NaI}\cdot 3\text{Me}_2\text{CO}$  are filtered off and, on warming in a flask, acetone distils off readily. [This method is more convenient than the one using the bisulfite addition compound.] It has also been purified by gas chromatography on a 20% free fatty acid phthalate (on Chromosorb P) column at  $100^\circ$ .

For efficiency of desiccants in drying acetone see Burfield and Smithers [*J Org Chem* **43** 3966 1978]. The water content of acetone can be determined by a modified Karl Fischer titration [Koupparis & Malmstadt *Anal Chem* **54** 1914 1982]. [*Beilstein* **1** IV 3180.]

**Rapid procedure:** Dry over anhydrous  $\text{CaSO}_4$  and distil.

**Acetone cyanohydrin** [75-86-5] **M 85.1, b  $48^\circ/2.5\text{mm}$ ,  $68-70^\circ/11\text{mm}$ ,  $78-82^\circ/15\text{mm}$ ,  $d_4^{20}$  0.93.** Dry the cyanohydrin with  $\text{Na}_2\text{SO}_4$ , and distil it as rapidly as possible under vacuum to avoid decomposition. Discard fractions boiling below  $78-82^\circ/15\text{mm}$ . Store it in the dark. **USE AN EFFICIENT FUME HOOD as HCN (POISONOUS) is always present.** [Cox & Stormont *Org Synth Coll.Vol. II* 7 1940, *Beilstein* **3** H 316, **3** IV 785.]

**Acetonedicarboxylic acid (3-oxoglutaric acid)** [542-05-2] **M 146.1, m  $138^\circ(\text{dec})$ ,  $\text{pK}^{25}$  3.10.** Crystallise it from ethyl acetate and store it over  $\text{P}_2\text{O}_5$ . It decarboxylates in hot water. [*Beilstein* **3** IV 1816.]

**Acetone semicarbazone** [110-20-3] **M 115.1, m  $187^\circ$ ,  $\text{pK}^{25}$  1.33.** Acetone semicarbazone crystallises from water or from aqueous EtOH. [*Beilstein* **3** H 101, **3** I 48, **3** II 81, **3** III 189, **3** IV 179.]

**Acetonitrile (methyl cyanide)** [75-05-8] **M 41.1, b  $81.6^\circ$ ,  $d^{25}$  0.77683,  $n_D^{20}$  1.3441,  $n^{25}$  1.34163.** Commercial acetonitrile is a by-product of the reaction of propylene and ammonia to acrylonitrile. The following procedure that significantly reduces the levels of acrylonitrile, allyl alcohol, acetone and \*benzene was used by Kiesel [*Anal Chem* **52** 2230 1988]. Methanol (300ml) is added to 3L of acetonitrile fractionated at high reflux ratio until the boiling temperature rises from  $64^\circ$  to  $80^\circ$ , and the distillate becomes optically clear down to  $\lambda = 240\text{nm}$ . Add sodium hydride (1g) free from paraffin, to the liquid, reflux for 10 minutes, and then distil rapidly until about 100ml of residue remains. Immediately pass the distillate through a column of acidic alumina, discarding the first 150ml of percolate. Add 5g of  $\text{CaH}_2$  and distil the first 50ml at a high reflux ratio. Discard this fraction, and collect the following main fraction. The best way of detecting impurities is by gas chromatography.

Usual contaminants in commercial acetonitrile include  $\text{H}_2\text{O}$ , acetamide,  $\text{NH}_4\text{OAc}$  and  $\text{NH}_3$ . Anhydrous  $\text{CaSO}_4$  and  $\text{CaCl}_2$  are inefficient drying agents. Preliminary treatment of acetonitrile with cold, saturated aqueous KOH is undesirable because of base-catalysed hydrolysis and the introduction of water. Drying by shaking with silica gel or Linde 4A molecular sieves removes most of the water in acetonitrile. Subsequent stirring with  $\text{CaH}_2$  until no further hydrogen is evolved leaves only traces of water and removes acetic acid. The acetonitrile is then fractionally distilled at high reflux, taking precaution to exclude moisture by refluxing over  $\text{CaH}_2$  [Coetzee *Pure Appl Chem* **13** 429 1966]. Alternatively, 0.5-1% (w/v)  $\text{P}_2\text{O}_5$  is often added to the distilling flask to remove most of the remaining water. Excess  $\text{P}_2\text{O}_5$  should be avoided because it leads to the formation of an orange polymer. Traces of  $\text{P}_2\text{O}_5$  can be removed by distilling from anhydrous  $\text{K}_2\text{CO}_3$ .

Kolthoff, Bruckenstein and Chantooni [*J Am Chem Soc* **83** 3297 1961] removed acetic acid from 3L of acetonitrile by shaking for 24 hours with 200g of freshly activated alumina (which had been reactivated by heating at  $250^\circ$  for 4 hours). The decanted solvent was again shaken with activated alumina, followed by five batches of 100-150g of anhydrous  $\text{CaCl}_2$ . (Water content of the solvent was then less than 0.2%.) It was shaken for 1 hour with 10g of  $\text{P}_2\text{O}_5$ , twice; and distilled in a  $1\text{m}\times 2\text{cm}$  column, packed with stainless steel wool and

protected from moisture by  $\text{CaCl}_2$  tubes. The middle fraction had a water content from 0.7 to 2mM. Traces of unsaturated nitriles can be removed by initially refluxing with a small amount of aqueous KOH (1ml of 1% solution per L). Acetonitrile can be dried by azeotropic distillation with dichloromethane, \*benzene or trichloroethylene. Isonitrile impurities can be removed by treatment with conc HCl until the odour of isonitrile has gone, followed by drying with  $\text{K}_2\text{CO}_3$  and distilling.

Acetonitrile is refluxed with, and distilled from alkaline  $\text{KMnO}_4$  and  $\text{KHSO}_4$ , followed by fractional distillation from  $\text{CaH}_2$ . (This is better than fractionation from molecular sieves or passage through a type H activated alumina column, or refluxing with  $\text{KBH}_4$  for 24 hours and fractional distillation) [Bell et al. *J Chem Soc, Faraday Trans 1* **73** 315 1977, Moore et al. *J Am Chem Soc* **108** 2257 1986].

Material suitable for polarography is obtained by refluxing over anhydrous  $\text{AlCl}_3$  (15g/L) for 1 hour, distilling, refluxing over  $\text{Li}_2\text{CO}_3$  (10g/L) for 1 hour and redistilling. It is then refluxed over  $\text{CaH}_2$  (2g/L) for 1 hour and fractionally distilled, retaining the middle portion. The product is not suitable for UV spectroscopy use. A better purification procedure uses refluxing over anhydrous  $\text{AlCl}_3$  (15g/L) for 1 hour, distilling, refluxing over alkaline  $\text{KMnO}_4$  (10g  $\text{KMnO}_4$ , 10g  $\text{Li}_2\text{CO}_3$ /L) for 15 minutes, and distilling. A further reflux for 1 hour over  $\text{KHSO}_4$  (15g/L), then distillation, is followed by refluxing over  $\text{CaH}_2$  (2g/L) for 1 hour, and fractional distillation. The product is protected from atmospheric moisture and stored under nitrogen [Walter & Ramalay *Anal Chem* **45** 165 1973]. Purification of "General Purity Reagent" for this purpose is not usually satisfactory because very large losses occur at the  $\text{KMnO}_4/\text{Li}_2\text{CO}_3$  step. For electrochemical work involving high oxidation fluorides, further reflux over  $\text{P}_2\text{O}_5$  (1g/ml for 0.5 hours) and distilling (discarding 3% of first and last fractions) and repeating this step is necessary. The distillate is kept over molecular sieves *in vacuo* after degassing, for 24 hours and distilling in a vacuum onto freshly activated 3A molecular sieves. The MeCN should have absorption at 200nm of  $<0.05$  ( $\text{H}_2\text{O}$  reference) and UV cutoff at *ca* 175nm. Also the working potential range of purified  $\text{Et}_4\text{N}^+ \text{BF}_4^-$  (0.1mol.dcm<sup>-3</sup> in the MeCN) should be +3.0 to -2.7V vs  $\text{Ag}^+/\text{Ag}^0$ . If these criteria are not realised then further impurities can be removed by treatment with activated neutral alumina (60 mesh) *in vacuo* before final molecular sieves treatment [Winfield *J Fluorine Chem* **25** 91 1984].

Acetonitrile has been distilled from  $\text{AgNO}_3$ , Payecting the middle fraction over freshly activated  $\text{Al}_2\text{O}_3$ . After standing for two days, the liquid is distilled from the activated  $\text{Al}_2\text{O}_3$ . The specific conductivity should be  $0.8\text{--}1.0 \times 10^{-8}$  mhos [Harkness & Daggett *Can J Chem* **43** 1215 1965]. Acetonitrile <sup>14</sup>C is best purified by gas chromatography and is water free, and distils at 81°. [Beilstein **2** H 183, **2** IV 419.]

**Rapid procedure:** Dry over anhydrous  $\text{K}_2\text{CO}_3$  for 24 hours, followed by further drying for 24 hours over 3A molecular sieves or boric anhydride, followed by distillation. Alternatively, stir over  $\text{P}_2\text{O}_5$  (5% w/v) for 24 hours then distil. However, this last method is not suitable for reactions with very acid sensitive compounds.

**Acetonylacetone (2,5-hexanedione)** [110-13-4] **M 114.2, m -9°, b 76-78°/13mm, 88°/25mm, 137°/150mm, 188°/atm, d<sub>4</sub><sup>20</sup> 0.9440, n<sub>D</sub><sup>20</sup> 1.423, pK<sup>25</sup> 18.7.** Purify it by dissolving in  $\text{Et}_2\text{O}$ , stirring with  $\text{K}_2\text{CO}_3$  (a quarter of the weight of dione), filtering, drying over anhydrous  $\text{Na}_2\text{SO}_4$  (**not**  $\text{CaCl}_2$ ), filtering again, evaporating the filtrate and distilling it in a vacuum. It is then redistilled through a 30cm Vigreux column (oil bath temperature 150°). It is miscible with  $\text{H}_2\text{O}$  and EtOH. The *dioxime* has **m** 137° (plates from  $\text{C}_6\text{H}_6$ ), the *mono-oxime* has **b** 130°/11mm, and the *2,4-dinitrophenylhydrazone* has **m** 210-212° (red needles from EtOH). It forms complexes with many metals. [Werner et al. *Chem Ber* **22** 2100 1989, for enol content see Gero *J Org Chem* **19** 1960 1954, *Beilstein* **1** IV 3688.]

**Acetoxime (acetone oxime)** [127-06-0] **M 73.1, m 63°, b 135°/760mm, d<sub>4</sub><sup>20</sup> 0.901, pK<sup>40</sup> 0.99.** It crystallises from petroleum ether (b 40-60°) and can be sublimed. [*Beilstein* **1** H 649, **1** IV 3202.]

**Acetoxyacetone (acetonyl acetate, acetol acetone)** [592-20-1] **M 116.1, b 65°/11mm, 73-75°/17mm, 174-176°/atm, d<sub>4</sub><sup>20</sup> 1.0757, n<sub>D</sub><sup>20</sup> 1.4141.** Distil it in a vacuum, then redistil it at atmospheric pressure. It is miscible with  $\text{H}_2\text{O}$ , but is slowly decomposed by it. Store it in a dry atmosphere. The *2,4-dinitrophenylhydrazone* has **m** 115-115.5° (from  $\text{CHCl}_3$ /hexane). [Perkin Jr *J Chem Soc* **59** 789 1891, Reich & Samuels *J Org Chem* **21** 68 1956, Nef *Justus Liebigs Ann Chem* **335** 260 1904, *Beilstein* **2** IV 297.]

**1-Acetoxy-1,3-butadiene (1,3-butadienyl acetate) cis-trans mixture** [1515-76-0] **M 112.1, b 42-43°/16mm, 51-52°/20mm, 60-61°/40mm, d<sub>4</sub><sup>20</sup> 0.9466, n<sub>D</sub><sup>20</sup> 1.4622.** The commercial sample is stabilised with 0.1% of *p*-

*tert*-butylcatechol. If the material contains crotonaldehyde (by IR, used in its synthesis), it should be dissolved in Et<sub>2</sub>O, shaken with 40% aqueous sodium bisulfite, then 5% aqueous Na<sub>2</sub>CO<sub>3</sub>, water, dried (Na<sub>2</sub>SO<sub>4</sub>) and distilled several times in a vacuum through a Widmer [*Helv Chim Acta* **7** 59 1924] or Vigreux column [Wichterle & Hudlicky *Col Czech Chem Commun* **12** 564 1947, Hagemeyer & Hull *Ind Eng Chem* **41** 2920 1949]. [*Beilstein* **2** III 295.]

**1-Acetoxy-2-butoxyethane (2-butyloxyethyl acetate)** [112-07-2] **M 160.2, b 61-62°/0.2mm, 75-76°/12mm, 185.5°/740mm, 188-192°/atm, d<sub>4</sub><sup>20</sup> 0.9425, n<sub>D</sub><sup>20</sup> 1.4121.** Shake the ester with anhydrous Na<sub>2</sub>CO<sub>3</sub>, filter and distil it in a vacuum. Redistillation can then be carried out at atmospheric pressure. [Dunbar & Bolstad *J Org Chem* **21** 1041 1956, *Beilstein* **2** IV 215.]

**2-Acetoxyethanol (2-hydroxyethyl acetate)** [542-59-6] **M 104.1, b 61°/1mm, 79-81°/12mm, 187°/761mm, d<sub>4</sub><sup>20</sup> 1.108, n<sub>D</sub><sup>20</sup> 1.42.** Dry the ester over K<sub>2</sub>CO<sub>3</sub> (not CaCl<sub>2</sub>), and distil it. [Davis & Ross *J Chem Soc* 3061 1950, rate of hydrolysis: Davis & Ross *J Chem Soc* 2706 1951, *Beilstein* **2** H 141, **2** I 66, **2** II 154, **2** III 303, **2** IV 214.]

**1-Acetoxy-2-ethoxyethane** [111-15-9] **M 132.2, b 30°/3mm, 49-50°/12mm, 156-159°, d<sub>4</sub><sup>20</sup> 0.97, n<sub>D</sub><sup>20</sup> 1.406.** Shake the ethoxy-ethane with anhydrous Na<sub>2</sub>CO<sub>3</sub>, filter and distil it in a vacuum. Redistillation can then be carried out at atmospheric pressure. [Dunbar & Bolstad *J Org Chem* **21** 1041 1956, *Beilstein* **2** IV 214.]

**1-Acetoxy-2-methoxyethane** [110-49-6] **M 118.1, b 30°/6mm, 40-41°/12mm, 140-144°/760mm, d<sub>4</sub><sup>20</sup> 1.009, n<sub>D</sub><sup>20</sup> 1.4011.** Shake the methoxy-ethane with anhydrous Na<sub>2</sub>CO<sub>3</sub>, filter and distil it in a vacuum. Redistillation can then be carried out at atmospheric pressure. [Dunbar & Bolstad *J Org Chem* **21** 1041 1956, *Beilstein* **2** IV 214.]

**S(-)-2-Acetoxypropionyl chloride** [36394-75-9] **M 150.6, b 51-53°/11mm, d<sub>4</sub><sup>20</sup> 1.19, n<sub>D</sub><sup>20</sup> 1.423, [α]<sub>D</sub><sup>27</sup> -33°, (c 4, CHCl<sub>3</sub>), [α]<sub>D</sub><sup>20</sup> -38° (c 4, CHCl<sub>3</sub>).** It is moisture sensitive and is hydrolysed to the corresponding acid. Check the IR spectrum. If the OH band above 3000cm<sup>-1</sup> is too large and broad then the mixture should be refluxed with pure acetyl chloride for 1 hour, evaporated and distilled under reduced pressure. [Julia & Sans *J Chromatographic Sci* **17** 651 1979, Dolittle & Heath *J Org Chem* **49** 5041 1984, *Beilstein* **3** II 189.]

**S-Acetoxy succinic anhydride** [59025-03-5] **M 158.1, m 58° (RS 81.5-82.5°, 86-87°), [α]<sub>D</sub><sup>20</sup> -26.0° (c 19, Me<sub>2</sub>CO), [α]<sub>D</sub><sup>20</sup> -28.4° (c 13, Ac<sub>2</sub>O).** Recrystallise it from Ac<sub>2</sub>O and dry it in a vacuum over KOH, or by washing it with dry Et<sub>2</sub>O due to its deliquescent nature. [Jones *J Chem Soc* 788 1933, Henrot et al. *Synth Commun* **16** 183 1986, Shiuuey et al. *J Org Chem* **52** 1040 1988, RS: Cohen et al. *J Am Chem Soc* **88** 5306 1966.]

**Acetylacetone (2,4-pentanedione, acac)** [123-54-6] **M 100.1, m -23°, 45°/30mm, b 140.4°/atm, d<sub>30.2</sub><sup>30.2</sup> 0.9630, n<sub>D</sub><sup>18.5</sup> 1.45178, pK<sub>1</sub><sup>25</sup> -5.0 (enol), -6.6 (keto), pK<sub>2</sub><sup>25</sup> 8.95.** Small amounts of acetic acid are removed by shaking with small portions of 2M NaOH until the aqueous phase remains faintly alkaline. The sample, after washing with water, is dried with anhydrous Na<sub>2</sub>SO<sub>4</sub>, and distilled through a modified Vigreux column [Cartledge *J Am Chem Soc* **73** 4416 1951]. An additional purification step is fractional crystallisation from the liquid. Alternatively, there is less loss of acetylacetone if it is dissolved in four volumes of \*benzene and the solution is shaken three times with an equal volume of distilled water (to extract acetic acid): the \*benzene is then removed by distillation at 43-53° and 20-30mm through a helices-packed column. It is then refluxed over P<sub>2</sub>O<sub>5</sub> (10g/L) and fractionally distilled under reduced pressure. The distillate (sp conductivity 4 × 10<sup>-8</sup> ohm<sup>-1</sup>cm<sup>-1</sup>) is suitable for polarography [Fujinaga & Lee *Talanta* **24** 395 1977]. To recover used acetylacetone, metal ions are stripped from the solution at pH 1 (using 100ml 0.1M H<sub>2</sub>SO<sub>4</sub>/L of acetylacetone). The acetylacetone is then washed with (1:10) ammonia solution (100ml/L) and with distilled water (100ml/L, twice), then treated as above. It complexes with Al, Be, Ca, Cd, Ce, Cu, Fe<sup>2+</sup>, Fe<sup>3+</sup>, Mn, Mg, Ni, Pb and Zn. [*Beilstein* **1** H 777, **1** I 401, **1** II 831, **1** III 3113, **1** IV 3662.]

**Acetyl bromide** [506-96-7] **M 123.0, b 76-77°, d<sub>4</sub><sup>20</sup> 1.65.** Boil acetyl bromide with PBr<sub>3</sub>/Ac<sub>2</sub>O for 1 hour, then distil the latter off and redistil it. Store it dry. [Burton & Degering *J Am Chem Soc* **62** 227 1940, *Beilstein* **2** IV

## 398.] LACHRYMATORY.

**2-Acetylbutyrolactone** [517-23-7] **M 128.1, b 105°/5mm, 120-123°/11mm, 142-143°/30mm, d<sub>4</sub><sup>20</sup> 1.1846, n<sub>D</sub><sup>20</sup> 1.459.** Purify the lactone by distillation, which will convert any free acid to the lactone, *alternatively* dissolve it in Et<sub>2</sub>O, wash well with 0.5N HCl, dry the organic layer and distil it. Its solubility in H<sub>2</sub>O is 20% v/v. The *2,4-dinitrophenylhydrazone* forms orange needles from MeOH, **m 146°**. The lactone hydrolyses in mineral acid to 2-acetyl-4-hydroxybutyric acid which can be converted to the *di-n-propylamine salt* with **m 68-70°**. The lactone is a **SKIN IRRITANT**. [Matsuhawa *Yakugaku Zasshi (J Pharm Soc Jpn)* **62** 417 1942, Willman & Schinz *Helv Chim Acta* **35** 2401 1952, *Beilstein* **17/11 V 16.**]

**Acetyl chloride** [75-36-5] **M 78.5, b 52°, d<sub>4</sub><sup>20</sup> 1.1051, n<sub>D</sub><sup>20</sup> 1.38976.** Reflux acetyl chloride with PCl<sub>5</sub> for several hours to remove traces of acetic acid, then distil it. Redistil it from one-tenth its volume of dimethylaniline or quinoline to remove free HCl. A.R. quality is freed from HCl by pumping it for 1 hour at -78° and distilling it into a trap at -196°. [*Beilstein* **2 IV 395.**] **LACHRYMATORY.**

**Acetylene** [74-86-2] **M 26.0, m -80.8°, b -84°, pK<sup>25</sup> ~25.** If very impure, acetylene should be purified by successive passage through spiral wash bottles containing, in this order, saturated aqueous NaHSO<sub>4</sub>, H<sub>2</sub>O, 0.2M iodine in aqueous KI (two bottles), sodium thiosulfate solution (two bottles), alkaline sodium hydrosulfite with sodium anthraquinone-2-sulfonate as indicator (two bottles), and 10% aqueous KOH solution (two bottles). The gas is then passed through a Dry-Ice trap and two drying tubes, the first containing CaCl<sub>2</sub>, and the second, Dehydrite [Mg(ClO<sub>4</sub>)<sub>2</sub>] [Conn et al. *J Am Chem Soc* **61** 1868 1939]. Acetone vapour can be removed from acetylene by passage through H<sub>2</sub>O, then conc H<sub>2</sub>SO<sub>4</sub>, or by passage through two gas traps at -65° and -80°, conc H<sub>2</sub>SO<sub>4</sub> and a soda lime tower, a tower of 1-mesh Al<sub>2</sub>O<sub>3</sub> then through H<sub>2</sub>SO<sub>4</sub> [Reichert & Nieuwland *Org Synth Coll Vol I* 229 1941, Wiley *Org Synth Coll Vol III* 853 1955, Jones & Whiting *Org Synth Coll Vol IV* 793 1963]. Sometimes it contains acetone and air. These can be removed by a series of bulb-to-bulb distillations, e.g. a train consisting of a conc H<sub>2</sub>SO<sub>4</sub> trap and a cold EtOH trap (-73°), or passage through H<sub>2</sub>O and H<sub>2</sub>SO<sub>4</sub>, then over KOH and CaCl<sub>2</sub>. [See Brandsma *Preparative Acetylenic Chemistry*, 1st Edn Elsevier 1971 p15, for pK, ISBN 0444409475, 2nd Edn Elsevier 1988, ISBN 0444429603, and below for sodium acetylide.] It is also available commercially as 10ppm in helium, and several concentrations in N<sub>2</sub> for instrument calibration. [*Beilstein* **1 IV 939.**]

**Sodium acetylide** [1066-26-8] **M 48.0,** is prepared by dissolving Na (23g) in liquid NH<sub>3</sub> (1L) and bubbling acetylene until the blue colour is discharged (ca 30 minutes) and evaporated to dryness [Saunders *Org Synth Coll Vol III* 416 1955], and is available commercially as a suspension in xylene/light mineral oil. [See entry in "Metal-organic Compounds", Chapter 5.]

**Acetylenedicarboxamide (Aquamycin, Cellocidin)** [543-21-5] **M 112.1, m 216-218°(dec), 219-221°(dec).** Acetylenedicarboxamide crystallises from MeOH and H<sub>2</sub>O [**m 190-192°(dec)** as *hemihydrate*]. When prepared from the ester + NH<sub>3</sub> it has **m 213°(dec)**. Also a melting point of 290-292°(dec) has been reported. [Saggimo *J Org Chem* **22** 1171 1857, Kharash et al. *J Org Chem* **10** 392 1945, Blomquist & Winslow *J Org Chem* **10** 156 1945, *Beilstein* **2 I 317, 2 III 1995, 2 IV 2295.**]

**Acetylenedicarboxylic acid (butynedioic acid)** [142-45-0] **M 114.1, m 179°(anhydrous), pK<sub>1</sub><sup>19</sup> 1.04, pK<sub>2</sub><sup>19</sup> 2.50.** The acid is soluble in Et<sub>2</sub>O and crystallises from Et<sub>2</sub>O/petroleum ether (**m 183-183.5°**), or H<sub>2</sub>O as the *dihydrate* which dehydrates in a desiccator over conc H<sub>2</sub>SO<sub>4</sub> in a vacuum. The *dipicrate* crystallises from aqueous ether. For the mono K salt see entry in "Metal-organic Compounds", Chapter 5. [Abbott et al. *Org Synth Coll Vol II* 10 1943, Huntress et al. *Org Synth Coll Vol IV* 329 1963, *Beilstein* **2 H 801, 2 I 317, 2 II 670, 2 III 1991, 2 IV 2290.**]

**N-Acetylenediamine** [1001-53-2] **M 102.1, m 50-51°, 51°, b 128°/3mm, 125-130°/5mm, 133-139°/27mm, pK<sup>25</sup> 9.28.** The acetyl-diamine has been fractionated under reduced pressure and fraction **b 125-130°/5mm** was refractionated, fraction **b 132-135°/4mm** was collected and solidified. It is a low melting *hygroscopic* solid which can be recrystallised from dioxane/Et<sub>2</sub>O. It is soluble in H<sub>2</sub>O, Et<sub>2</sub>O and \*C<sub>6</sub>H<sub>6</sub>. The *p-toluenesulfonate salt* can be recrystallised from EtOH/EtOAc (1:8), has **m 125-126°**, but the free base cannot be recovered from it by basifying and extracting with CH<sub>2</sub>Cl<sub>2</sub>. The *picrate* has **m 175°** (from EtOH) [Aspinall *J*

*Am Chem Soc* **63** 853 1941, Hall *J Am Chem Soc* **78** 2570 1956]. [*Beilstein* **4** IV 1193.]

**Acetyl fluoride** [557-99-3] **M 62.0, b 20.5°/760mm,  $d_4^{20}$  1.032.** Purify acetyl fluoride by fractional distillation. It attacks glass and is sold in steel cylinders. [*Beilstein* **2** H 172, **2** I 79, **2** II 175, **2** III 385, **2** IV 393.] **TOXIC and LACHRYMATORY.**

**Acetyl iodide** [507-02-8] **M 170.0, b 39.5-40°/64mm, 108°/760mm.** Purify it by fractional distillation. [*Beilstein* **2** H 174, **2** I 80, **2** II 177, **2** III 393, **2** IV 399.] **TOXIC and LACHRYMATORY.**

**3-(S-Acetylmercapto)isobutyric acid** [RS 33325-40-5] **M 162.2, m 40-40.5°, b ca 120°/1.25mm,  $pK_{Est}$  ~4.0.** Distil the acid under high vacuum and recrystallise it from \*C<sub>6</sub>H<sub>6</sub>. [Fredga & Mastersson *Chem Abstr* **38** 3616 1944.]

**Acetyl methanesulfonate** [5539-53-7] **M 170.2, b <120°/ <0.01mm.** The main impurity is methanesulfonic acid. Reflux it with redistilled acetyl chloride for 6-10 hours, i.e. until no further HCl is absorbed in a trap, and exclude moisture. Distil off excess of AcCl and carefully distil it below 0.001mm with the bath temperature below 120° to give the anhydride as a pale yellow oil which solidifies below 0°. Below ~130° it gives the disulfonic anhydride, and above ~130° polymers are formed, and it is used for cleaving ethers [Preparation, IR, NMR: Karger & Mazur *J Org Chem* **36** 528, 532 1971]. [*Beilstein* **2** H 166, **2** III 349.]

**3-(Acetylthio)propionic acid** [41345-70-4] **M 148.2, m 48-52°, 52-54°, b 127-128°/3mm,  $pK_{Est}$  ~4.2.** Purify the propionic acid by distillation in a vacuum. It has  $\lambda_{max}$  at 231nm ( $\epsilon$  4200). It is a potential enzyme inhibitor [Noda et al. *J Am Chem Soc* **75** 914 1953, Clegg et al. *J Am Chem Soc* **121** 5319 2004, Clegg & Hutchinson *Angew Chem, Int Edn* **43** 3716 2005]. [*Beilstein* **3** III 551, **3** IV 731.]

**N-Acetylthiourea** [591-08-2] **M 118.2, m 164-165°, 166-168°.** Recrystallise the thiourea from AcOH; the solid is washed with Et<sub>2</sub>O and dried in air then at 100°. [Zahradnik *Col Czech Chem Commun* **24** 3678 1959, *Beilstein* **3** IV 354.]

**cis-Aconitic acid** [585-84-2] **M 174.1, m 126-129°(dec).** Crystallise the *cis*-acid from water by cooling (solubility is 1g in 2ml of water at 25°). Dry it in a vacuum desiccator. [*Beilstein* **2** IV 2405.]

**trans-Aconitic acid (1,2,3-propenetriscarboxylic acid)** [4023-65-8] **M 174.1, m 195°(dec), m 198-199°(dec), 204-205°(dec),  $pK_1^{25}$  2.81,  $pK_2^{25}$  4.46.** Purify the *trans*-acid by dissolving it in AcOH (77g/150ml), filtering and cooling. The acid separates (55g) as colourless needles. A further quantity (10g) can be obtained by reducing the volume of the filtrate. The acid is dried in air then in a vacuum desiccator over NaOH. The acid can be recrystallised from Me<sub>2</sub>CO/CHCl<sub>3</sub>. The highest melting point is obtained with the very dry acid. A melting point of 209° was obtained on a Dennis bar [Dennis & Shalton *J Am Chem Soc* **52** 3128 1930, Bruce *Org Synth Coll Vol II* 12 1943]. [*Beilstein* **2** IV 2405.]

**cis-Aconitic anhydride** [6318-55-4] **M 156.1, m 75°, 76-78°, 78-78.5°.** Reflux it in xylene (7.5 parts) for 1 hour, then evaporate and recrystallise the residue from \*C<sub>6</sub>H<sub>6</sub>. Alternatively, reflux it in Ac<sub>2</sub>O, evaporate and recrystallise from \*C<sub>6</sub>H<sub>6</sub>. It is sensitive to moisture, store dry. [IR: Groth & Dahlén *Acta Chem Scand* **21** 291 1967, Malachowski & Maslowski *Chem Ber* **61** 2523 1928, NMR: Gawron & Mahajan *Biochemistry* **5** 2335 1966.] [*Beilstein* **18/8** V 530.]

**Acrolein (acraldehyde, 2-propenal)** [107-02-8] **M 56.1, fp -86.95°, b 25°/100mm, 52.69°/760mm (dt/dp 0.0355°/mm,  $d_4^{20}$  0.839,  $n_D^{25}$  1.3992.** Purify acrolein by fractional distillation, under nitrogen, drying with anhydrous CaSO<sub>4</sub> and then distilling under vacuum. Blacet, Young and Roof [*J Am Chem Soc* **59** 608 1937] distilled it under nitrogen through a 90cm column packed with glass rings. To avoid formation of diacryl, the vapour is passed through an ice-cooled condenser into a receiver cooled in an ice-salt mixture and containing 0.5g catechol. The acrolein is then distilled twice from anhydrous CuSO<sub>4</sub> at low pressure, catechol being placed in the distilling flask and the receiver to avoid polymerisation. [Alternatively, hydroquinone (1% of the final

solution) can be used.] **Respiratory irritant, work in an efficient fume cupboard.** [*Beilstein* 1 IV 3435.]

**Acrolein diacetyl acetal (1,1-diacetoxy-2-propene).** [869-29-4] **M 158.2, b 75°/10mm, 184°/atm,  $d_4^{20}$  1.08,  $n_D^{20}$  1.4203.** Check the NMR spectrum. If it is not satisfactory, then add  $\text{Ac}_2\text{O}$  and a drop of conc  $\text{H}_2\text{SO}_4$  and heat at 50° for 10 minutes. Then add anhydrous  $\text{NaOAc}$  (ca 3g/100g of liquid) and fractionate. Note that it forms an azeotrope with  $\text{H}_2\text{O}$ , so do not add  $\text{H}_2\text{O}$  at any time. It is a **highly flammable and TOXIC** liquid; use protective gloves. [Smith et al. *J Am Chem Soc* 73 5282 1951, *Beilstein* 2 H 154, 2 I 72, 2 III 356, 2 IV 291.]

**Acrolein diethyl acetal (3,3-diethoxy-1-propene or 1,1-diethoxy-2-propene)** [3054-95-3] **M 130.2, b 120-125°/atm,  $n_D^{20}$  1.398-1.407.** Add  $\text{Na}_2\text{CO}_3$  (ca 3.5%) and distil it using an efficient column, or better use a spinning band column. [Witzemann et al. *Org Synth Coll Vol II* 17 1943, *Beilstein* 1 H 727, 1 I 378, 1 III 2960, 1 IV 3437.]

**Acrolein dimethyl acetal (1,1-dimethoxy-2-propene)** [6044-68-4] **M 102.1, b 87.5-88°/750mm, 89-90°/760mm,  $d_4^{20}$  0.86,  $n_D^{20}$  1.3962.** Fractionally distil it (after adding 0.5g of hydroquinone) under reduced pressure through an all glass column (40 cm × 2.5 cm) packed with glass helices and provided with a heated jacket and a total reflux variable take-off head. Stainless steel Lessing rings (1/8 x 1/8 in) or gauze have also been used as packing. It is a **highly flammable and TOXIC** liquid; keep away from the skin. [Hall & Stern *J Chem Soc* 2657 1955, *Beilstein* 1 IV 3437.]

**Acrolein semicarbazone** [6055-71-6] **M 113.1, m 171°.** It crystallises from water in needles. [Auwers & Heineke *Justus Liebigs Ann Chem* 458 202 1927, *Beilstein* 1 II 785.]

**Acrylamide** [79-06-1] **M 71.1, m 84°, b 125°/25mm.** Crystallise acrylamide from acetone, chloroform, ethyl acetate, methanol or \*benzene/chloroform mixture, then vacuum dry and store it in the dark under vacuum. Recrystallise it from  $\text{CHCl}_3$  by dissolving 200g in 1L, heating to boiling and filtering without suction in a warmed funnel through Whatman 541 filter paper; allowing to cool to room temperature and keeping at -15° overnight. The crystals are collected with suction in a cooled funnel and washed with 300ml of cold MeOH. The crystals are air-dried in a warm oven. [Dawson et al. *Data for Biochemical Research*, Oxford Press 1986 p. 449, *Beilstein* 2 IV 1471.]

**CAUTION:** *Acrylamide is extremely TOXIC (neurotoxic), and precautions must be taken to avoid skin contact or inhalation. Use gloves and handle in a well-ventilated fume cupboard.*

**Acrylic acid** [79-10-7] **M 72.1, m 13°, b 30°/3mm, 70°/50mm,  $d_4^{20}$  1.051,  $\rho_K^{25}$  4.25.** It can be purified by steam distillation, or vacuum distillation through a column packed with copper gauze to inhibit polymerisation. (This treatment also removes inhibitors such as methylene blue that may be present.) Azeotropic distillation of the water with \*benzene converts aqueous acrylic acid to the anhydrous material. [*Beilstein* 2 H 397, 2 I 186, 2 II 383, 2 III 1215, 2 IV 1455.]

**Acrylonitrile** [107-13-1] **M 53.1, b 78°,  $d_4^{20}$  0.806,  $n_D^{25}$  1.3886.** Wash acrylonitrile with dilute  $\text{H}_2\text{SO}_4$  or dilute  $\text{H}_3\text{PO}_4$ , then with dilute  $\text{Na}_2\text{CO}_3$  and water. Dry it with  $\text{Na}_2\text{SO}_4$ ,  $\text{CaCl}_2$  or (better) by shaking with molecular sieves. Fractionally distil it under  $\text{N}_2$ . It can be stabilised by adding 10ppm *tert*-butyl catechol. Immediately before use, the stabiliser can be removed by passage through a column of activated alumina (or by washing with 1% NaOH solution if traces of water are permissible in the final material), followed by distillation. Alternatively, shake it with 10% (w/v) NaOH to extract inhibitor, and then wash it in turn with 10%  $\text{H}_2\text{SO}_4$ , 20%  $\text{Na}_2\text{CO}_3$  and distilled water. Dry for 24 hours over  $\text{CaCl}_2$  and fractionally distil under  $\text{N}_2$  taking fraction boiling at 75.0-75.5°C (at 734mm). Store it with 10ppm *tert*-butyl catechol. Acrylonitrile is distilled off when required. [Burton et al. *J Chem Soc, Faraday Trans 1* 75 1050 1979, *Beilstein* 2 IV 1473.]

**Acryloyl chloride** [814-68-6] **M 90.5, b 72-74°/740mm, 74°/760mm,  $d_4^{20}$  1.1127,  $n_D^{20}$  1.4337.** Distil acryloyl chloride rapidly through an efficient 25cm column after adding 0.5g of hydroquinone/200g of chloride, and then redistil it carefully at atmospheric pressure preferably in a stream of dry  $\text{N}_2$ . [Stempel et al. *J Am Chem Soc* 72 72 1950, Gresham et al. *J Am Chem Soc* 72 2299 1950, *Beilstein* 2 IV 1471.] **The liquid is an irritant and is**

## TOXIC.

**Adipic acid** [124-04-9] **M 146.1**, **m 154-154.5°**, **b 159.5°/0.1mm**, **191°/5mm**, **205.5°/10mm**, **222.5°/20mm**, **337.5°/760mm**, **pK<sub>1</sub><sup>20</sup> 4.42**, **pK<sub>2</sub><sup>20</sup> 5.41**; **pK<sub>1</sub><sup>25</sup> 4.44**, **pK<sub>2</sub><sup>25</sup> 5.45**; **pK<sub>1</sub><sup>40</sup> 4.54**, **pK<sub>2</sub><sup>40</sup> 5.59**. For use as a volumetric standard, adipic acid is crystallised once from hot water with the addition of a little animal charcoal, dried at 120° for 2 hours, then recrystallised from acetone and again dried at 120° for 2 hours. Other purification procedures include crystallisation from ethyl acetate and from acetone/petroleum ether, fusion, followed by filtration and crystallisation from the melt, and preliminary distillation under vacuum. [*Beilstein* 2 H 649, 2 I 277, 2 II 572, 2 III 1705, 2 IV 1956.]

**Diethyl adipate** [141-28-6] **M 202.3**, **m -20° to -19°**, **b 111°/5mm**, **125°/10mm**, **133°/15mm**, **251°/atm**, **d<sup>25</sup> 1.0034**, **n<sub>D</sub><sup>20</sup> 1.42776**, is prepared in the usual way by boiling the acid and excess EtOH in the presence of a catalytic amount of H<sub>2</sub>SO<sub>4</sub> and distilled. [*Beilstein* 2 H 652, 2 I 277, 2 II 574, 2 III 1721, 2 IV 1960.]

**Adiponitrile (1,4-dicyanobutane)** [111-69-3] **M 108.14**, **m 2.4°**, **b 123°/0.5mm**, **153°/6mm**, **175°/26mm**, **184°/30mm**, **295°/atm**, **d<sub>4</sub><sup>20</sup> 0.9396**, **n<sub>D</sub><sup>20</sup> 1.4371**. Reflux adiponitrile over P<sub>2</sub>O<sub>5</sub> and POCl<sub>3</sub>, and fractionally distil it, then fractionate it through an efficient column. **The liquid is TOXIC and is an IRRITANT.** [Braun & Rudolph *Chem Ber* 67 1770 1934, Reppe et al. *Justus Liebigs Ann Chem* 596 127 1955, Gagnon et al. *Can J Chem* 34 1662 1956, Copley et al. *J Am Chem Soc* 62 228 1940, *Beilstein* 2 IV 1947.]

**Agaricic acid [1-(*n*-hexadecyl)citric acid]** [666-99-9] **M 416.6**, **m 142°(dec)**, **[α]<sub>D</sub> -9.8° (in NaOH)**, **pK<sub>Est(1)</sub> ~2.7**, **pK<sub>Est(2)</sub> ~4.2**, **pK<sub>Est(3)</sub> ~5.5**. Crystallise the acid from EtOH. The *trihydrazide* has **m 170°(dec)** (from EtOH). [Brandänge et al. *Acta Chem Scand B* 31 307 1977, *Beilstein* 3 I 186, 3 II 372, 3 III 1109, 3 IV 1284.]

**Agmatine sulfate [5-guanidinopent-1-ylamine sulfate]** [2482-00-0] **M 228.3**, **m 231°**, **pK<sub>Est(1)</sub> ~9.1**, **pK<sub>Est(2)</sub> ~13.0**. Crystallise the salt from aqueous MeOH. The *free base* has **m 101.5-103°**, the *gold chloride hydrochloride* crystallises from H<sub>2</sub>O with **m 223°(dec)**, and the *picrate* has **m 236-238°**. [Odo *J Chem Soc Jpn* 67 132 1946, *Beilstein* 4 I 420, 4 II 703, 4 III 575, 4 IV 1291.]

**Aldol (3-hydroxybutanal)** [107-89-1] **M 88.1**, **b 80-81°/20mm**, **d<sup>16</sup> 1.109**. An ethereal solution of aldol is washed with a saturated aqueous solution of NaHCO<sub>3</sub>, then with water. The non-aqueous layer is dried with anhydrous CaCl<sub>2</sub> and distilled immediately before use. The fraction, **b 80-81°/20mm**, is collected as a thick liquid which decomposes at 85°/atm. It is a sedative and a hypnotic, but is used in perfumery. [Mason et al. *J Am Chem Soc* 76 2255 1954]. [*Beilstein* 1 H 824, 1 I 419, 1 II 868, 1 III 3195, 1 IV 3984.]

**Aleuritic acid [*RS*-erythro-9,10,16-trihydroxyhexadecanoic acid]** [533-87-9] **M 304.4**, **m 100-101°**, **pK<sub>Est</sub> ~4.9**. Crystallise this *RS*-acid from aqueous EtOH. It is soluble in MeOH, and forms a less soluble crystalline sodium salt. The *methyl ester* **m 72-73°**, **b 235°/0.2mm**, is best prepared by reaction with diazomethane and forms fine feathery needles; it is soluble in MeOH, EtOH, CHCl<sub>3</sub>, Me<sub>2</sub>CO, slightly soluble in \*C<sub>6</sub>H<sub>6</sub> and insoluble in petroleum ether. The *ethyl ester* [6003-09-4] **m 59°**, crystallises in needles from EtOH. The *hydrazide* [6003-10-7] crystallises from EtOH and has **m 139-140°**.

The *RS*-acid has been isolated from Shellac although it has two asymmetric carbon atoms, and possibly contains the *RS*-erythro or *RS*-cis form [Gidvani *J Chem Soc* 306 1944, Sengupta & Bose *J Sci Ind Res (India)* 11B 458 1952]. A stereoisomer have been synthesised [Hunsdieker *Chem Ber* 76 142 1943, Hunsdieker *Chem Ber* 77 185 1944]. [*Beilstein* 3 III 901.]

***n*-Alkylammonium chloride *n*=2,4,6.** Recrystallise them from EtOH or an EtOH/Et<sub>2</sub>O mixture. [Hashimoto & Thomas *J Am Chem Soc* 107 4655 1985, Chu & Thomas *J Am Chem Soc* 108 6270 1986.]

***n*-Alkyltrimethylammonium bromide *n*=10,12,16.** Recrystallise them from an EtOH/Et<sub>2</sub>O mixture. [Hashimoto & Thomas *J Am Chem Soc* 107 4655 1985.]

**Allene (propadiene)** [463-49-0] **M 40.1**, **m -146°**, **b -32°**. Freeze allene in liquid nitrogen, evacuate, then thaw out. This cycle is repeated several times, then the allene is frozen in a methylcyclohexane/liquid nitrogen bath and pumped for some time. It has also been purified by HPLC. [Cripps & Kiefer *Org Synth* 42 12 1962,

*Beilstein* 1 IV 966.]

**neo-Alloocimene (allocimene B, *tc*-2,6-dimethyl-2,4,6-octatriene)** [7216-56-0; *cis/trans* mixture 673-84-7; *trans/trans* 3016-19-1] **M 136.2, b 80°/13mm, 196-198°/atm,  $d_4^{20}$  0.8161,  $n_D^{20}$  1.5437.** Fractionally distil allocimene through an efficient column and repeatedly distil it at 15mm through a long column of glass helices, with a final distillation from sodium under nitrogen. It should be stabilised with *ca* 0.1% of hydroquinone. Its UV has  $\lambda_{\max}$  nm( $\epsilon$  M<sup>-1</sup>cm<sup>-1</sup>) at 290 (32 500), 279 (41,900) and 278 (42,870). [Alder et al. *Justus Liebigs Ann Chem* 609 1 1957, O'Connor & Goldblatt *Anal Chem* 26 1726 1954, *Beilstein* 1 IV 1106.]

**Allyl acetate** [591-87-7] **M 100.1, b 103°,  $d_4^{20}$  0.928,  $n_D^{17}$  1.4004,  $n_D^{20}$  1.4040.** The ester is freed from peroxides by standing with crystalline ferrous ammonium sulfate, then washed with 5% NaHCO<sub>3</sub>, followed by saturated CaCl<sub>2</sub> solution. Dry it with Na<sub>2</sub>SO<sub>4</sub> and fractionally distil it in an all-glass apparatus. **FLAMMABLE LIQUID.** [*Beilstein* 2 H 136, 2 IV 180.]

**Allylacetic acid (pent-4-enoic acid)** [591-80-0] **M 100.1, m -22.5°, b 83-84°/12mm, 90°/15mm, 187-189°/~760mm,  $d_4^{20}$  0.9877,  $n_D^{20}$  1.4280,  $pK^{25}$  4.68.** Distil the acid through an efficient column (allyl alcohol has **b** 95-97°). It is characterised as the *S*-benzylisothiuronium salt **m** 155-158° (from 96% EtOH, or aqueous EtOH) [Friediger & Pedersen *Acta Chem Scand* 9 1425 1955], and the 4-bromophenacyl ester has **m** 59.5-60.5° (from 90% EtOH). Its solubility at 18° in solvents is: pyridine (57%), AcOH (7.3%), MeOH (5.4%), Me<sub>2</sub>CO (3.2%), MeOAc (2.8%), EtOH (5.4%), H<sub>2</sub>O (1.8%), PrOH (1.6%), isoPrOH (0.27%). [Brown & Berkowski *J Am Chem Soc* 74 1894 1952, *Beilstein* 2 IV 1542.]

**Allyl alcohol** [107-18-6] **M 58.1, b 98°,  $d_4^{20}$  0.857,  $n_D^{20}$  1.4134.** It can be dried with K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, or by azeotropic distillation with \*benzene followed by distillation under nitrogen. It is difficult to obtain it free of peroxide. It has also been refluxed with magnesium and fractionally distilled [Hands & Norman *Ind Chem* 21 307 1945]. [*Beilstein* 1 IV 2079.]

**Allylamine** [107-11-9] **M 57.1, b 52.9°,  $d_4^{20}$  0.761,  $n_D^{20}$  1.42051,  $pK^{25}$  9.49.** Purify allylamine by fractional distillation from calcium chloride. It causes sneezing and tears. [*Beilstein* 4 IV 1057.]

**Allyl bromide** [106-95-6] **M 121, b 70°,  $d_4^{20}$  1.398,  $n_D^{20}$  1.46924.** Wash the bromide with NaHCO<sub>3</sub> solution then distilled water, dry (CaCl<sub>2</sub> or MgSO<sub>4</sub>), and fractionally distil. Protect it from strong light. [*Beilstein* 1 IV 754.] **LACHRYMATORY, HIGHLY TOXIC and FLAMMABLE.**

**Allyl butyl ether** [3739-64-8] **M 114.2, b 64-65°/120mm, 117.8-118°/763mm,  $d_4^{20}$  1.4057,  $n_D^{20}$  0.7829.** Check the IR for the presence of OH str vibrations; if so then wash it well with H<sub>2</sub>O, dry it with CaCl<sub>2</sub> and distil it through a good fractionating column. **The liquid is an irritant.** [Watanabe et al. *J Org Chem* 23 1666 1958, Schueler & Hanna *J Am Chem Soc* 73 3528 1951, *Beilstein* 1 IV 2084.]

**Allyl chloride** [107-05-1] **M 76.5, b 45.1°,  $d_4^{20}$  0.939,  $n_D^{20}$  1.4130.** Likely impurities include 2-chloropropene, propyl chloride, *iso*-propyl chloride, 3,3-dichloropropane, 1,2-dichloropropane and 1,3-dichloropropane. Purify it by washing with conc HCl, then with Na<sub>2</sub>CO<sub>3</sub> solution, dry it with CaCl<sub>2</sub>, and distil it through an efficient column [Oae & Vanderwerf *J Am Chem Soc* 75 2724 1953]. [*Beilstein* 1 IV 738.] **LACHRYMATORY, TOXIC.**

**Allyl chloroformate** [2937-50-0] **M 120.5, b 56°/97mm, 109-110°/atm,  $d_4^{20}$  1.14,  $n_D^{20}$  1.4223.** Wash the chloroformate several times with cold H<sub>2</sub>O to remove alcohol and HCl and dry it over CaCl<sub>2</sub>. It is **important** to dry well before distilling *in vacuo*. Note that the receiver should be cooled in ice to avoid loss of distillate into the trap and vacuum pump. The liquid is **highly TOXIC and flammable.** [Fierz-David & Müller *J Chem Soc* 125 26 1924, Strain et al. *J Am Chem Soc* 72 1254 1950, *Beilstein* 3 IV 29.]

**Allyl cyanide (3-butene nitrile)** [109-75-1] **M 67.1, b -19.6°/1.0mm, 2.9°/5mm, 14.1°/5mm, 26.6°/20mm, 48.8°/60mm, 60.2°/100mm, 98°/400mm, 119°/760mm,  $d_4^{20}$  0.8341,  $n_D^{20}$  1.406.** The nitrile should be redistilled at atmospheric pressure, then distilled under a vacuum to remove the final traces of HCN from the

residue. Note that the residue from the first distillation may be difficult to remove from the flask and should be treated with conc  $\text{HNO}_3$  then  $\text{H}_2\text{O}$  and finally hot  $\text{EtOH}$  (**CARE**). Allyl cyanide has an onion-like odour and is stable to heat. It forms a complex with  $\text{AlCl}_3$  (2:2)  $m$   $41^\circ$ , and (3:2)  $m$   $120^\circ$ . **All operations should be done in an efficient fume hood as the liquid is flammable, may contain cyanide and is HIGHLY TOXIC.** [Supniewski et al. *Org Synth Coll Vol I* 46 1941, *Beilstein* 2 IV 1491.]

**Allyl disulfide (diallyl disulfide)** [2179-57-9]  $M$  146.3,  $b$   $58\text{--}59^\circ/5\text{mm}$ ,  $79\text{--}81^\circ/20\text{mm}$ ,  $138\text{--}139^\circ/\text{atm}$ ,  $d_4^{20}$  1.01,  $n_D^{20}$  1.541. Purify the disulfide by fractional distillation until the molar refractivity is in uniformly good agreement with the calculated value [Small et al. *J Am Chem Soc* 69 1710 1947]. It has also been purified by gas chromatography [retention times: Carson & Wong *J Org Chem* 24 175 1959, UV: Koch *J Chem Soc* 395 1949]. It is present in garlic. [*Beilstein* 1 IV 2098.]

**Allyl iodide (3-iodopropene)** [556-56-9]  $M$  167.7,  $b$   $103^\circ/760\text{mm}$ ,  $d^{12}$  1.848. Purify allyl iodide in a dark room by washing with aqueous  $\text{Na}_2\text{SO}_3$  to remove free iodine, then dry with  $\text{MgSO}_4$  and distil at  $43^\circ/90$  mm or at atmospheric pressure to give a very pale yellow liquid. (This material, dissolved in hexane, can be stored in a light-protected tight container at  $-5^\circ$  for up to three months before free iodine can be detected, by its colour in the solution.) Store it away from light. [Sibbett & Noyes *J Am Chem Soc* 75 761 1953, *Beilstein* 1 H 202, 1 I 84, 1 II 172, 1 III 714, 1 IV 761.]

**Allylisocyanate** [1476-23-9]  $M$  83.1,  $b$   $84^\circ/\text{atm}$ ,  $87\text{--}89^\circ/\text{atm}$ ,  $d_4^{20}$  0.94,  $n_D^{20}$  1.417. Purify it as for allylisothiocyanate below and it is **TOXIC**. [*Beilstein* 4 IV 1081.]

**Allylisothiocyanate** [57-06-7]  $M$  99.2,  $m$   $-80^\circ$ ,  $b$   $84\text{--}85^\circ/80\text{mm}$ ,  $150^\circ/760\text{mm}$ ,  $151^\circ/\text{atm}$ ,  $d_4^{20}$  1.017,  $n_D^{20}$  1.5268. Fractionate the isothiocyanate using an efficient column, preferably in a vacuum. It is a yellow **pungent, irritating and TOXIC (suspected CARCINOGEN) liquid**. Store it in a sealed tube under  $\text{N}_2$ . The *N'*-benzylthiourea derivative has  $m$   $94.5^\circ$  (from aqueous  $\text{EtOH}$ ) [Weller et al. *J Am Chem Soc* 74 1104 1952]. [*Beilstein* 4 IV 1081.]

***N*-Allylthiourea (thiosinamine)** [109-57-9]  $M$  116.2,  $m$   $70\text{--}73^\circ$ ,  $78^\circ$ . Recrystallise it from  $\text{H}_2\text{O}$ . It is soluble in 30 parts of cold  $\text{H}_2\text{O}$ , and it is soluble in  $\text{EtOH}$  but insoluble in  $^*\text{C}_6\text{H}_6$ . It has also been recrystallised from acetone,  $\text{EtOH}$  or ethyl acetate, after decolorising with charcoal. The white crystals have a bitter taste with a slight garlic odour and are **TOXIC**. An unstable crystalline form is obtained by recrystallising from the melt. [McCrone et al. *Anal Chem* 21 421 1949, *Beilstein* 4 IV 1072.]

***N*-Allylurea** [557-11-9]  $M$  100.1,  $m$   $85^\circ$ . It crystallises from  $\text{EtOH}$ ,  $\text{EtOH}/\text{ether}$ ,  $\text{EtOH}/\text{chloroform}$  or  $\text{EtOH}/\text{toluene}$ . [*Beilstein* 4 IV 1070.]

**Aminoacetaldehyde dimethyl acetal (2,2-dimethoxyethylamine)** [22483-09-6]  $M$  105.1,  $m$   $<-78^\circ$ ,  $b$   $139.5^\circ/768\text{mm}$ ,  $137\text{--}139^\circ/\text{atm}$ ,  $d_4^{20}$  0.9676  $n_D^{20}$  1.4144. Dry the acetal over  $\text{KOH}$  pellets and distil it through a 30cm vacuum jacketed Vigreux column. [Lawson *J Am Chem Soc* 75 3398 1953, Erickson et al. *J Am Chem Soc* 77 6640 1955, *Beilstein* 4 IV 1918.]

**Aminoacetonitrile bisulfate** [151-63-3]  $M$  154.1,  $m$   $125^\circ(\text{dec})$ ,  $\text{pK}^{25}$  5.34 ( $\text{NH}_2$ ). Recrystallise the hydrogensulfate (1:1) from  $\text{EtOH}/\text{Et}_2\text{O}$  (hygroscopic leaflets). The **Sulfate (2:1)** [5466-22-8] crystallises as flat prisms from  $\text{H}_2\text{O}/\text{EtOH}$  with  $m$   $166^\circ(\text{dec})$ . [Stephen *J Chem Soc* 871 1931, Anslow & King *J Chem Soc* 2465 1929, *Beilstein* 4 III 1120.]

**Aminoacetonitrile hydrochloride** [6011-14-9]  $M$  92.5,  $m$   $166\text{--}167^\circ$ ,  $172\text{--}174^\circ$ ,  $\text{pK}^{25}$  5.24 ( $\text{NH}_2$ ). The salt recrystallises from dilute  $\text{EtOH}$  as *hygroscopic* leaflets. It is best to crystallise it from absolute  $\text{EtOH}/\text{Et}_2\text{O}$  (1:1) and then recrystallise it from absolute  $\text{EtOH}$ . The melting point recorded ranges from  $144$  to  $174^\circ$ . The *free base* has  $b$   $58^\circ/15\text{mm}$  with partial decomposition. [Klages *J Prakt Chem* [2] 65 189 1902, Mange *J Am Chem Soc* 56 2197 1934, Goldberg & Kelly *J Chem Soc* 1371 1947, *Beilstein* 4 H 344, 4 I 468, 4 II 783, 4 III 1120, 4 IV 2363.]

**2-Amino-1-butanol** [*RS*(±) 96-20-8, *R*( ) 5856-63-3, *S*(+) 5856-62-2] **M 89.1**, **m** ~ -2°, **b** 78-80°/10mm and 179-183°/atm for (±), 172-174°/atm for (+) or ( ), **pK<sup>0</sup> 10.353**, **pK<sup>20</sup> 9.672**, **pK<sup>60</sup> 8.555**. They are purified by shaking with solid NaOH, filtering and distilling through a short column. The *oxalate salt* of the racemate has **m** 176°. They are strong bases and should be stored under N<sub>2</sub> in the absence of CO<sub>2</sub>. The *enantiomers* have [ $\alpha$ ]<sub>D</sub><sup>20</sup> +12.5° and -12.5° (c 2, EtOH). [Johnson & Degering *J Org Chem* **8** 7 1943, Nagao et al. *J Org Chem* **51** 2392 1986, Santaniello et al. *J Chem Soc, Perkin Trans 1* 919 1985, *Beilstein 4* H 291, **4** IV 1705.]

**2-Aminoethanol (ethanolamine)** [141-43-5] **M 61.1**, **f** 10.5°, **b** 72-73°/12mm, 171.1°/760mm, **d**<sub>4</sub><sup>20</sup> 1.012, **n**<sub>D</sub><sup>20</sup> 1.14539, **pK<sup>25</sup> 9.51**. It decomposes slightly when distilled at atmospheric pressure, with the formation of conducting impurities. Fractional distillation at about 12mm pressure is most satisfactory. After distillation, 2-aminoethanol is further purified by repeated washing with ether and crystallising from EtOH (at low temperature). After fractional distillation in the absence of CO<sub>2</sub>, it is twice crystallised by cooling, followed again by distillation. It is *hygroscopic*, and absorbs CO<sub>2</sub> from the atmosphere. [Reitmeier et al. *J Am Chem Soc* **62** 1943 1940.] It can be dried by azeotropic distillation with dry \*benzene. [*Beilstein 4* IV 1406.]

**2-Aminoethanol hydrochloride** [2002-24-6] **M 97.6**, **m** 75-77°. Recrystallise the salt from EtOH. It is deliquescent; store it dry. [*Beilstein 4* IV 1406.]

**2-Aminoethyl hydrogen sulfate (sulfuric acid mono-2-aminoethyl ester)** [926-39-6] **M 141.1**, **m** 285-287° (chars at 275°). Crystallise the sulfate ester from water or dissolve it in water and add EtOH. Wash this with Et<sub>2</sub>O and dry it *in vacuo*. [*Beilstein 4* III 1414.]

**S-(2-Aminoethyl)isothiuronium bromide hydrobromide** [56-10-0] **M 281.0**, **m** 194-195°. Crystallise the salt from absolute EtOH/ethyl acetate or MeOH. Store dry as it is *hygroscopic* in a humid atmosphere. It is a radioprotective agent. When refluxed in EtOH for 16 hours or H<sub>2</sub>O for 30 minutes, it decomposes to 2-amino-4(5H)-thiazoline hydrobromide which on recrystallisation from isoPrOH/EtOAc has **m** 175-176° [Doherty et al. *J Am Chem Soc* **79** 5667 1957].

**(2-Aminoethyl)trimethylammonium chloride hydrochloride (chloramine chloride hydrochloride)** [3399-67-5] **M 175.1**, **m** 268°(dec). Crystallise the hydrochloride from EtOH. The material is very soluble in H<sub>2</sub>O. [*Beilstein 4* II 690.]

**Aminomalononitrile toluene-4-sulfonate** [5098-14-6] **M 253.4**, **m** 168-170°, 172°(dec), **pK<sub>Est</sub> ~ 1.3**. It forms colourless crystals on recrystallisation from MeCN (1.8g in 100mL) using activated charcoal. Wash the crystals with dry Et<sub>2</sub>O and dry them at 25°/1mm. Recovery is ~80%. [Ferris et al. *Org Synth Coll Vol V* 32 1973.]

**(±)-2-Amino-4-methylhexane (Forthane, 1,3-dimethylpentylamine)** [105-41-9] **M 115.2**, **b** 131-133°/atm, 130-135°/atm, 135-136°/atm, **d**<sub>4</sub><sup>20</sup> 0.760, **n**<sub>D</sub><sup>15</sup> 1.4160, **n**<sub>D</sub><sup>24</sup> 1.4160, **pK<sup>25</sup> 10.54**. This strong base is obtained by catalytic hydrogenation of 4-methylhexan-2-one oxime (57g) in EtOH (50ml) with Raney Ni (6g) in a bomb at 75-80° and 1000psi of H<sub>2</sub>, cool, filter off the catalyst, acidify with HCl, and evaporate to dryness. Basify the residue with aqueous NaOH, extract it with Et<sub>2</sub>O, dry the extract (MgSO<sub>4</sub>), filter, evaporate, and distil to give the base in 75-80% yield. It forms a *sulfate salt* **m** 215-220°(dec), and it readily forms a *carbonate salt*, hence it should be stored in the absence of CO<sub>2</sub>. It is physiologically active as an opium pressor and is adrenergic. [Rohrmann & Shonle *J Am Chem Soc* **66** 1516 1944, Chiang *J Clin Chem Soc* **18** 65 1951, USP to E. Lilly 2350318 (1943) *Chem Abstr* **39** 1510 1945, USP 2386273 (1943) *Chem Abstr* **40** 598 1946, *Beilstein 4* III 378, **4** IV 747.]

**2-Amino-2-methyl-1,3-propanediol** [115-69-5] **M 105.1**, **m** 111°, **b** 151-152°/10mm, **pK<sup>25</sup> 8.80**. Crystallise the diol three times from MeOH, dry in a stream of dry N<sub>2</sub> at room temperature, then in a vacuum oven at 55°. Store it over CaCl<sub>2</sub> [Hetzer & Bates *J Phys Chem* **66** 308 1962]. [*Beilstein 4* IV 1881.]

**2-Amino-2-methyl-1-propanol** ( $\beta$ -aminoisobutanol) [124-68-5] **M 89.4, m 24°, 31°, b 67°/10mm, 164-166°/760mm,  $d_4^{20}$  0.935,  $n_D^{20}$  1.45,  $pK^{25}$  9.71.** Purify it by distilling and fractional freezing. The *hydrochloride* [3207-12-3] has **m 204°-206°.** [Beilstein 4 III 783, 4 IV 1740.]

***n*-Amyl acetate** (*n*-pentyl acetate) [628-63-7] **M 130.2, b 149.2°,  $d_4^{20}$  0.876,  $n_D^{20}$  1.40228.** Shake the ester with saturated NaHCO<sub>3</sub> solution until neutral, washed it with water, dry with MgSO<sub>4</sub> and distil it. The ester has also been purified by repeated fractional distillation through an efficient column or spinning band column. [Timmermann & Hennant-Roland *J Chim Phys* 52 223 1955, Mumford & Phillips *J Chem Soc* 75 1950, <sup>1</sup>H NMR: Crawford & Foster *Can J Phys* 34 653 1956, Beilstein 2 IV 152.]

***n*-Amyl alcohol** (1-pentanol) [71-41-0] **M 88.2, b 138.1°,  $d^{15}$  0.818,  $n_D^{20}$  1.4100.** Dry 1-pentanol with anhydrous K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, filter and fractionally distil it. It has also been treated with 1-2% of sodium and heated at reflux for 15 hours to remove water and chlorides. Traces of water can be removed from the near-dry alcohol by refluxing it with a small amount of sodium in the presence of 2-3% *n*-amyl phthalate or succinate followed by distillation (see *ethanol*).

Small amounts of amyl alcohol have been purified by esterifying with *p*-hydroxybenzoic acid, recrystallising the ester from CS<sub>2</sub>, saponifying with ethanolic-KOH, drying with CaSO<sub>4</sub> and fractionally distilling [Olivier *Rec Trav Chim Pays Bas* 55 1027 1936]. [Beilstein 1 IV 1640.]

***tert*-Amyl alcohol** (2-methyl-2-butanol) [75-85-4] **M 88.2, m -12°, b 102.3°,  $d^{15}$  0.8135,  $n_D^{20}$  1.4058.** Reflux it with K<sub>2</sub>CO<sub>3</sub>, CaH<sub>2</sub>, CaO or sodium, then fractionally distil. The near-dry alcohol is further dried by refluxing with Mg activated with iodine, as described for *ethanol*. Further purification is possible using fractional crystallisation and zone refining at <-10° or preparative gas chromatography. [Beilstein 1 IV 1668.]

***n*-Amylamine** [1-aminopentane] [110-58-7] **M 87.2, b 105°,  $d_4^{20}$  0.752,  $pK^{25}$  10.63.** Dry it by prolonged shaking with NaOH pellets, then distilling. Store it in a CO<sub>2</sub>-free atmosphere. [Beilstein 4 IV 674.]

***n*-Amyl bromide** (*n*-pentylbromide) [110-53-2] **M 151.1, b 129.7°,  $d_4^{20}$  1.218,  $n_D^{20}$  1.445.** Wash the bromide with conc H<sub>2</sub>SO<sub>4</sub>, then water, 10% Na<sub>2</sub>CO<sub>3</sub> solution, again with water, dry with CaCl<sub>2</sub> or K<sub>2</sub>CO<sub>3</sub>, and fractionally distil it just before use. [Beilstein 1 IV 312.]

***n*-Amyl chloride** (1-chloropentane) [543-59-9] **M 106.6, b 107.8°,  $d_4^{20}$  0.882,  $n_D^{20}$  1.41177.** Purify as for *sec*-amyl chloride. [Beilstein 1 IV 309.]

***sec*-Amyl chloride** (1-chloro-2-methylbutane) [616-13-7] **M 106.6, b 52.2°/150mm, 96-97° (100°)/760mm,  $d_4^{20}$  0.886,  $n_D^{20}$  1.412.** Purify the chloride by stirring vigorously with 95% H<sub>2</sub>SO<sub>4</sub>, replacing the acid when it becomes coloured, until the layer remains colourless after 12 hours stirring. The amyl chloride is then washed with saturated Na<sub>2</sub>CO<sub>3</sub> solution, then distilled water, and dried with anhydrous MgSO<sub>4</sub>, followed by filtration, and distillation through a 10-in Vigreux column. *Alternatively*, a stream of oxygen containing 5% ozone is passed through the amyl chloride for three times as long as it takes to cause the first coloration of starch iodide paper by the exit gas. The liquid is washed with NaHCO<sub>3</sub> solution to hydrolyse the ozonides and remove organic acids prior to drying and fractional distillation [Chien & Willard *J Am Chem Soc* 75 6160 1953]. The *S*(+)-*enantiomer* has **b 50-51°/140mm, 100°/760, mm,  $[\alpha]_D^{20}$  +1.64° (neat)** [Brown et al. *J Am Chem Soc* 62 3437 1940]. [Beilstein 1 H 134, 1 I 46, 1 III 356, 1 IV 326.]

***tert*-Amyl chloride** (2-chloro-2-methylbutane) [594-36-5] **M 106.6, b 86°,  $d_4^{20}$  0.866.** Methods of purification commonly used for other alkyl chlorides lead to decomposition. Unsaturated contaminants are removed by chlorination with a small amount of chlorine in bright light, followed by distillation [Chien & Willard *J Am Chem Soc* 75 6160 1953]. [Beilstein 1 H 134, 1 I 46, 1 II 100, 1 III 357, 1 IV 324.]

**Amylene** ( $\beta$ -*iso*-amylene, 2-methyl-2-butene) [513-35-9] **M 70.1, b 37-38°/~760mm,  $d_4^{20}$  0.663,  $n_D^{20}$  1.387.** Distil amylenes and collect the distillate at low temperature. It has also been distilled from sodium. **FLAMMABLE.** It is available in steel cylinders and has a short shelf life. [Beilstein 1 H 211, 1 I 87, 1 II 187, 1 III 788, 1 IV 820.]

**Amyl ether (dipentyl ether)** [693-65-2] M 158.3, b 186.8°,  $d_4^{20}$  0.785,  $n_D^{20}$  1.41195. Repeatedly reflux amyl ether over sodium and distil it. [Beilstein 1 IV 1643.]

**Arachidic (eicosanoic C<sub>20</sub>) acid** [506-30-9] M 312.5, m 77°,  $pK_{Est}$  ~5.0. Crystallise the C<sub>20</sub> acid from absolute EtOH. [Beilstein 2 IV 1276.]

**Arachidic alcohol (1-eicosanol C<sub>20</sub>)** [629-96-9] M 298.6, m 65.5° (71°), b 200°/3mm. Crystallise the C<sub>20</sub> alcohol from \*benzene or \*benzene/petroleum ether. [Beilstein 1 IV 1900.]

**Azelaic acid (1,9-nonanedioic acid, heptane-1,7-dicarboxylic acid)** [123-99-9] M 188.2, m 105-106°, b 225°/10mm, 256°/50mm,  $pK_1^{25}$  4.53,  $pK_2^{25}$  5.33. Recrystallise it from H<sub>2</sub>O (charcoal) or thiophene-free \*benzene. The acid can be dried by azeotropic distillation with toluene, the residual toluene solution is then cooled and filtered, and the precipitate is dried in a vacuum oven. It has been purified by zone refining or by sublimation onto a cold finger at 10<sup>-3</sup>torr. It distils above 360° with partial formation of the anhydride. The *dimethyl ester* has m -3.9° and b 140°/8mm. [Hill & McEwen *Org Synth Coll Vol II* 53 1943, Beilstein 2 IV 2055.]

**2,2'-Azobis(isobutyronitrile) (α,α'-bis-azo-[2-methylpropionitrile], AIBN)** [78-61-1] M 164.2, m 102.1-103.2°, 103°, 103-104°. Crystallise the nitrile from Et<sub>2</sub>O, Me<sub>2</sub>CO, CHCl<sub>3</sub>, aqueous EtOH or MeOH. It has also been crystallised from absolute EtOH below 40° in subdued light. Dry it under vacuum at room temperature over P<sub>2</sub>O<sub>5</sub> and store it under vacuum in the dark at <-10° until required. Also crystallise it from CHCl<sub>3</sub> solution by addition of petroleum ether (b <40°). It is a radical inhibitor. [Askham et al. *J Am Chem Soc* 107 7423 1985, Ennis et al. *J Chem Soc, Dalton Trans* 2485 1986, Inoue & Anson *J Phys Chem* 91 1519 1987, Tanner *J Org Chem* 52 2142 1987.] It is prepared by the oxidation of *N,N'*-bis(isobutyronitrile)hydrazine by addition of the solid hydrazine (4.15g, 0.025mole) to a solution of HNO<sub>3</sub> (100%, 6.5g, 0.1mol) and Ac<sub>2</sub>O (30ml) (0.1mole) which are previously mixed at -30°. As the temperature rises to -20° the solid begins to dissolve and the colour of the solution turns to green, and dissolution is complete at ~5°. After stirring at 25° for 0.5 hours, the solution is poured onto cracked ice and H<sub>2</sub>O (100ml), the product separates out, is filtered off, washed free from acid with H<sub>2</sub>O. The solid (m 95-100°) is dried *in vacuo*, recrystallised from 50% EtOH and from Et<sub>2</sub>O to give pure AIBN (2.5g, 60%), m 105°. [Picard & Boivin *Can J Chem* 29 223 1951.] Alternatively, the hydrazine (0.989g, 6mmol) is added to a solution of glacial acetic acid (40ml) and concentrated H<sub>2</sub>SO<sub>4</sub> (3ml) at ~0°, followed by addition of NaMnO<sub>2</sub>·3H<sub>2</sub>O (1.8g) in H<sub>2</sub>O (25ml). MnO<sub>2</sub> separates immediately, excess of NaMnO<sub>2</sub> and MnO<sub>2</sub> are reduced with aqueous sodium bisulfite and poured into H<sub>2</sub>O. AIBN separates, is filtered off, washed with H<sub>2</sub>O, dried in a vacuum desiccator and recrystallised as above to give pure *azo* compound (0.61g, 63%). [Overberger & Lebovitz *J Am Chem Soc* 76 2722 1954, Overberger et al. *J Am Chem Soc* 71 2661 1949.] [Beilstein 4 H 563, 4 I 566, 4 III 1750, 4 IV 3377.]

**Azomethane (dimethyldiimide)** [503-28-6] M 58.1, m -78°, b 1.5°,  $d_0^{15}$  0.981,  $n_D^{19}$  1.3933. Purify azomethane by distillation in a vacuum line and store it in the dark at -80°. It is soluble in EtOH, Et<sub>2</sub>O and EtOAc. It can be **EXPLOSIVE**. [Beilstein 4 H 562, 4 I 566, 4 II 966, 4 III 1747, 4 IV 3366.]

**B.A.L. (British Anti-Lewesite)** see 2,3-dimercapto-1-propanol.

**Batyl alcohol (rac-3-[1-octadecyloxy]-1,2-propanediol)** [544-62-7] M 344.6, m 70.5-71°. Batyl alcohol crystallises from aqueous Me<sub>2</sub>CO, EtOH or petroleum ether (b 40-60°). [Taguchi & Armarego *Med Res Rev* 18 pp43-88 1998, Beilstein 1 IV 2758.]

**Behenoyl chloride (docosanoyl chloride)** [21132-76-3] M 359.0, m 40°. If the IR shows OH bands, then it should be dissolved in oxalyl chloride in \*C<sub>6</sub>H<sub>6</sub> solution and warmed at 35° for 24 hours in the absence of moisture, evaporated and distilled in a vacuum of 10<sup>-5</sup>mm. It is soluble in \*C<sub>6</sub>H<sub>6</sub> and Et<sub>2</sub>O. It is moisture sensitive and is **LACHRYMATORY**. [Francis et al. *J Chem Soc* 1001 1937, Levene & Taylor *J Biol Chem* 59 905 1924, Beilstein 2 III 1076.]

**Biacetyl (butan-2,3-dione)** [431-03-8] **M 86.1, b 88°, d<sub>4</sub><sup>20</sup> 0.981, n<sub>D</sub><sup>18.5</sup> 1.3933.** Dry biacetyl over anhydrous CaSO<sub>4</sub>, CaCl<sub>2</sub> or MgSO<sub>4</sub>, then distil it in a vacuum under nitrogen, taking the middle fraction and storing it at Dry-Ice temperature in the dark (to prevent polymerisation). [*Beilstein* 1 IV 3644.]

**Biguanide** [56-03-1] **M 101.1, m 130° pK<sub>1</sub><sup>25</sup> 3.1, pK<sub>2</sub><sup>25</sup> 12.8.** Crystallise biguanide from EtOH. It gives a red Cu derivative, and it forms salts with many metals. The *monohydrochloride* has **m 235°** [38664-03-8] and the *dihydrochloride* forms plates with **m 248°** (213-214°, also reported) [25836-74-2]. [*Beilstein* 3 H 93, 3 I 44, 3 II 76, 3 III 171, 3 IV 162.]

**Bis-acrylamide (N,N'-methylene bisacrylamide)** [110-26-9] **M 154.2, m >300°.** Recrystallise the amide from MeOH (100g dissolved in 500ml boiling MeOH) and filter without suction in a warmed funnel. Allow to stand at room temperature and then at -15°C overnight. The crystals are collected with suction in a cooled funnel and washed with cold MeOH. The crystals are air-dried in a warm oven. [*Beilstein* 2 IV 1472.] **VERY TOXIC (neurotoxic).**

**Bis(β-chloroethyl)amine hydrochloride** [821-48-7] **M 178.5, m 214-215°, 216-217°, pK<sub>Est</sub> ~5.8 (free base).** Crystallise the salt from Me<sub>2</sub>CO or MeOH/Et<sub>2</sub>O. The *picrate* has **m 112-113°** (from EtOH or Me<sub>2</sub>CO). [*Mann J Chem Soc* 464 1934, *Ward J Am Chem Soc* 57 915 1935, *Beilstein* 4 III 238.]

**Bis(β-chloroethyl) ether** [111-44-4] **M 143.0, b 94°/33mm, 178.8°, d<sub>4</sub><sup>20</sup> 1.220, n<sub>D</sub><sup>20</sup> 1.45750.** Wash the ether with conc H<sub>2</sub>SO<sub>4</sub>, then Na<sub>2</sub>CO<sub>3</sub> solution, dry with anhydrous Na<sub>2</sub>CO<sub>3</sub>, and finally pass it through a 50cm column of activated alumina before distillation. *Alternatively*, wash it with 10% ferrous sulfate solution to remove peroxides, then H<sub>2</sub>O, dry with CaSO<sub>4</sub>, and distil it in a vacuum. Add 0.2% of catechol to stabilise it. [*Beilstein* 1 IV 1375, *Kamm & Waldo J Am Chem Soc* 43 2223 1921.] **VERY TOXIC.**

**2,2'-Bis[di-(carboxymethyl)-amino]diethyl ether, (HOOCCH<sub>2</sub>)<sub>2</sub>NCH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>N-(CH<sub>2</sub>COOH)<sub>2</sub>** [923-73-9] **M 336.3, pK<sub>1</sub><sup>20</sup> 1.8, pK<sub>2</sub><sup>20</sup> 2.76, pK<sub>3</sub><sup>20</sup> 8.84, K<sub>4</sub><sup>20</sup> 9.47.** Crystallise it from EtOH.

**Bis(2-ethylhexyl) sebacate** [‘dioctyl’ sebacate, di(2-ethylhexyl) 1,8-octanedicarboxylate] [122-62-3] **M 426.7, b 212°/1mm, d<sub>4</sub><sup>25</sup> 1.914, n<sub>D</sub><sup>20</sup> 1.4496.** If it is acidic due to hydrolysis (effervesces with NaHCO<sub>3</sub>), then purify by dissolving in Et<sub>2</sub>O, shaking with aqueous Na<sub>2</sub>CO<sub>3</sub>, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and distil the residue at a high vacuum. Note that 2-ethylhexan-1-ol has **b 184-185°/760mm** with estimated **b ~40-50°/1mm** (see Figure 1, and Table 2A, Chapter 1) and will distil at a much lower temperature than the ester. Otherwise it should be distilled through a short column under high vacuum and the higher boiling fraction is redistilled. This vacuum pump oil can be prepared by esterifying decanedioic acid with 2-ethylhexanoic acid by Fischer-Speier’s method (Fischer & Speier *Chem Ber* 28 1150 1895), whereby dry HCl gas is bubbled through the alcohol until its weight is increased by ~10%, the acid is added, the mixture is heated at 100° until esterification is complete (~2-3 hours), and purified as above. [*Bruno USPat* 2628249 1953, *GBPat* 747260 1956, *Beilstein* 2 IV 2803.]

**N,N-Bis(2-hydroxyethyl)-2-aminoethanesulfonic acid (BES)** [10191-18-1] **M 213.3, m 150-155°, pK<sub>25</sub> 7.17.** Crystallise BES from aqueous EtOH. [*Beilstein* 4 IV 3290.]

**Bis(2-hydroxyethyl)amino-tris-(hydroxymethyl)methane (BIS-TRIS)** [6976-37-0] **M 209.2, m 89°, 104°, pK<sub>20</sub> 6.46.** Crystallise BIS-TRIS from hot 1-butanol and dry it in a vacuum at 25°.

**N,N-Bis(2-hydroxyethyl)glycine (BICINE)** See in “Amino Acids and Peptides”, Chapter 7.

**Bis(2-mercaptoethyl)sulfone (BMS)** [145626-87-5] **M 186.3, m 57-58°, pK<sub>1</sub><sup>25</sup> 7.9, pK<sub>2</sub><sup>25</sup> 9.0.** BMS recrystallises from hexane as white fluffy crystals. Large amounts are best recrystallised from de-oxygenated H<sub>2</sub>O (charcoal). It is a good alternative reducing agent to dithiothreitol. Its IR (film) has  $\nu_{\max}$  □□ 2995, 2657, 1306, 1248, 1124 and 729 cm<sup>-1</sup>. The synthetic intermediate *thioacetate* has **m 82-83°** (white crystals from CCl<sub>4</sub>). The *disulfide* is purified by flash chromatography on SiO<sub>2</sub> and elution with 50% EtOAc/hexane, recrystallised from hexane, and has **m 137-139°**. [*Lamoureux & Whitesides J Org Chem* 58 633 1993, *Beilstein* 1 IV 2455.]

**Bis(trichloromethyl) carbonate (triphosgene)** [32315-10-9] M 296.8, m 79-83°, 81-83°, b 203-206°(slight dec). It is a good solid substitute for phosgene (using a third mol per mol). Crystallise it from petroleum ether (b 60-80°), wash it with anhydrous cold Et<sub>2</sub>O, de-gas it at 200mm then dry at 0.1mm (over H<sub>2</sub>SO<sub>4</sub>). It has IR with  $\nu_{\max}$  at 900 and 1900 cm<sup>-1</sup>. It is a **lachrymator**, is **TOXIC**, and should be handled with gloves and in an efficient fume hood. [Hales et al. *J Chem Soc* 620 1957, Eckert & Forster *Angew Chem, Int Ed Engl* 26 894 1987, *Aldrichimica Acta* 21 47 1988, *Beilstein* 3 H 17, 3 I 8, 3 II 16, 3 III 36, 3 IV 33.]

**Bistrifluoroacetamide (BTFA)** [407-24-9] M 209.1, m 85°, b 135-136°/744mm, 141°/760mm. A major impurity is trifluoroacetamide. Add trifluoroacetic anhydride to BTFA, reflux for 2 hours and fractionate using a Vigreux column at atmospheric pressure. [Donike *J Chromatogr* 78 273 1973, *Beilstein* 2 IV 471.]

**Biuret (allophanic acid amide, carbamoylurea)** [108-19-0] M 103.1, sinters at 218° and chars at 270°, pK<sub>1</sub><sup>25</sup> -0.88, pK<sub>2</sub><sup>25</sup> >4. Crystallise biuret from EtOH. [*Beilstein* 3 IV 141.]

**N-Bromoacetamide** [79-15-2] M 138.0, m 102-105°, 107-109°, 108°(anhydrous). A possible contaminant is CH<sub>3</sub>CONBr<sub>2</sub>. Recrystallise it from CHCl<sub>3</sub>/hexane (1:1, seed if necessary) or water and dry over CaCl<sub>2</sub>. It is a brominating agent. [Oliveto & Gerold *Org Synth Coll Vol IV* 104 1963.] Alternatively, dissolve it in the minimum volume of warm H<sub>2</sub>O (60°), then cool in an ice bath, collect the crystals and dry them in an anhydrous atmosphere, dissolve in Et<sub>2</sub>O, chill and evaporate till crystallisation. Dry the crystals *in vacuo* at 25°, then at 45° (m 108°). Crystallise from CHCl<sub>3</sub> (m 103°). Estimate the available Br iodometrically [Buckles et al. *J Org Chem* 23 483 1958]. [*Beilstein* 2 H 181, 2 I 82, 2 II 180, 2 III 406, 2 IV 417.]

**Bromoacetic acid** [79-08-3] M 138.9, m 50°, b 118°/15mm, 208°/760mm, d<sup>25</sup> 1.93, pK<sup>25</sup> 2.92. Crystallise bromoacetic acid from pet ether (b 40-60°). A diethyl ether solution of it is passed through an alumina column, and the ether is evaporated at room temperature under vacuum. It is best obtained by distillation from a Claisen (flask immersed in an oil bath) fitted with an insulated Vigreux column and the fraction b 108-110°/30mm is collected. It is light and moisture sensitive. [Natelson & Gottfried *Org Synth Coll Vol III* 381 1955, *Beilstein* 2 IV 526.] **LACHRYMATORY and is a skin IRRITANT.**

**Bromoacetone** [598-31-2] M 137.0, b 31.5°/8mm, 63.5-64°/50mm, 137°/atm, d<sup>23</sup> 1.643. Stand bromoacetone over anhydrous CaCO<sub>3</sub>, filter, distil it under low vacuum, and store it with CaCO<sub>3</sub> in the dark at 0°. [Levene *Org Synth Coll Vol II* 88 1943.] **Violently LACHRYMATORY and skin IRRITANT.**

**2-Bromobutane** [78-76-2] M 137.0, b 91.2°, d<sub>4</sub><sup>20</sup> 1.255, n<sub>D</sub><sup>20</sup> 1.4367, n<sup>25</sup> 1.4341. Wash 2-bromobutane with conc HCl, water, 10% aqueous NaHSO<sub>3</sub>, and then water. Dry it with CaCl<sub>2</sub>, Na<sub>2</sub>SO<sub>4</sub> or anhydrous K<sub>2</sub>CO<sub>3</sub>, and fractionally distil it through a 1m glass helices packed column. [*Beilstein* 1 IV 261.]

**Bromoform** [75-25-2] M 252.8, m 8.1°, 55-56°/35mm, 149.6°/760mm, d<sup>15</sup> 2.9038, d<sup>30</sup> 2.86460, n<sup>15</sup> 1.60053, n<sub>D</sub><sup>20</sup> 1.5988. The storage and stability of bromoform and chloroform are similar. Ethanol, added as a stabiliser, is removed by washing with H<sub>2</sub>O or with saturated CaCl<sub>2</sub> solution, and the CHBr<sub>3</sub>, after drying with CaCl<sub>2</sub> or K<sub>2</sub>CO<sub>3</sub>, is fractionally distilled. Prior to distillation, CHBr<sub>3</sub> has also been washed with conc H<sub>2</sub>SO<sub>4</sub> until the acid layer is no longer coloured, then dilute NaOH or NaHCO<sub>3</sub>, and H<sub>2</sub>O. A further purification step is fractional crystallisation by partial freezing. [*Beilstein* 1 IV 82.]

**(±)-2-Bromohexadecanoic acid (2-bromopalmitic acid)** [18263-25-7] M 335.3, m 51-53°, 52.3-52.5°, 53°, pK<sub>Est</sub> ~3.2. Recrystallise the acid from petroleum ether (b 60-80°, charcoal) and finally from EtOH. The *ethyl ester* has b 177-178°/2mm, d<sub>28</sub><sup>28</sup> 1.0484, n<sub>D</sub><sup>20</sup> 1.4560. [IR: Sweet & Estes *J Org Chem* 21 1426 1956, *Beilstein* 2 IV 1184.]

**6-Bromohexanoic acid (6-bromocaproic acid)** [4224-70-8] M 195.1, m 32-33°, 35°, 34-36°, b 129-130°/5mm. It has been prepared by the oxidation of 6-bromohexanol with concentrated HNO<sub>3</sub> (sp. gr. 1.42, one hour for addition at 25-30°, stir at ~25° for 4 hours then at 100° for 45 minutes) [Degering & Boatright *J Am Chem Soc* 72 5137]. It is made more conveniently by oxidation of cyclohexanone (174 g) with Caro's acid (using a mixture prepared from 919g of potassium persulfate for 10-15 minutes as in Barger et al. *J Chem Soc*

718 1937), and the crude lactones (200g) are treated with a cooled mixture of 48% HBr (1L) and concentrated H<sub>2</sub>SO<sub>4</sub> (240ml). After standing at ~25° for 2 hours, then for 4 hours at 100°, the mixture is cooled, poured into H<sub>2</sub>O, the organic layer is separated, the aqueous layer is saturated with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, and is extracted with Et<sub>2</sub>O. The combined organic layers and extracts are washed with saturated (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, dried (Mg<sub>2</sub>SO<sub>4</sub>), evaporated and distilled in a vacuum. [See Brown & Partridge *J Am Chem Soc* **66** 839 1944, Degering & Boatright *J Am Chem Soc* **72** 5137 1950]. [Beilstein **2** II 287, **2** III 737, **2** IV 940.] When the acid is refluxed with SOCl<sub>2</sub> in the absence of moisture until all the HCl and SO<sub>2</sub> have evolved, then evaporated, and the residue is distilled in a vacuum, it provides **6-bromohexanoyl chloride** [22809-37-6] **M 213.5, b 130°/20mm, d<sup>25</sup> 1.395, n<sub>D</sub><sup>20</sup> 1.486** as an almost colourless oil which is typically reactive and corrosive like most acid chlorides. [Osmond *J Chem Soc* 3469 1951]. *6-Bromohexanamide* has **m 107-108°**, and the *anilide* has **m 84°**. [Beilstein **2** IV 940.]

**S-(+)-1-Bromo-2-methylbutane** [534-00-9] **M 151.1, b 38.2°/39mm, 49°/62mm, 60.8°(57-58°)/100mm, 65-65.6°/140mm, 116-122°/atm, d<sub>4</sub><sup>20</sup> 1.2232, n<sub>D</sub><sup>20</sup> 1.4453, [α]<sub>D</sub><sup>20</sup> +5.1° (c 5, CHCl<sub>3</sub>) (neat, +5.8°)**. Wash the bromobutane with ice-cold H<sub>2</sub>O, dry by freezing, shake it twice with an equal volume of H<sub>2</sub>SO<sub>4</sub> at 0°, and twice with an equal volume of H<sub>2</sub>O at 0°. Freeze-dry and keep over freshly heated (and then cooled) K<sub>2</sub>CO<sub>3</sub>, and distil it through a vacuum jacketed column of broken glass. *Alternatively*, dissolve it in petroleum ether (b 40-60°), wash it with 5% NaOH, conc H<sub>2</sub>SO<sub>4</sub> (at 0°), then H<sub>2</sub>O, dry (CaCl<sub>2</sub>), evaporate it and distil. [Heller *J Am Chem Soc* **74** 4858 1952, Foley *J Am Chem Soc* **81** 2779 1959, Easton & Hargreaves *J Chem Soc* 1413 1959, Crombie & Harper *J Chem Soc* 2685 1950, Beilstein **1** IV 327.]

**2-Bromo-2-methylpropane** [507-19-7] **M 137.0, b 71-73°, d<sub>4</sub><sup>20</sup> 1.218, n<sub>D</sub><sup>20</sup> 1.429**. Neutralise the bromomethylpropane with K<sub>2</sub>CO<sub>3</sub>, distil, and dry it using molecular sieves (5A), then distil it in a vacuum and degas it by the freeze-pump-thaw technique. Seal it under vacuum. [Beilstein **1** IV 295.]

**1-Bromooctadecane** [112-89-0] **M 333.4, m 26°, 27.3°, 28-30°, b 178-179°/2mm, 214-218°/15mm, d<sub>4</sub><sup>20</sup> 0.976, n<sub>D</sub><sup>20</sup> 1.461**. Twice recrystallise bromooctadecane from the melt, then distil it under vacuum three times taking the middle cut. *Alternatively*, wash the oil with aqueous Na<sub>2</sub>SO<sub>4</sub>, then conc H<sub>2</sub>SO<sub>4</sub> (cool) and again with aqueous Na<sub>2</sub>SO<sub>4</sub>, and then fractionally distil it. [Meyer & Ried *J Am Chem Soc* **55** 1574 1933, Hoffmann & Smyth *J Am Chem Soc* **72** 171 1950, IR: LeFèvre et al. *Aust J Chem* **12** 743 1959, IR: Brini-Fritz *Bull Soc Chim Fr* 516 1957, Beilstein **1** IV 555.]

**(±)-2-Bromopentane** [107-81-3] **M 151.1, b 117.2°/753mm, 116-117°/atm, 117.5°/740mm, d<sub>4</sub><sup>20</sup> 1.2190, n<sub>D</sub><sup>20</sup> 1.4401**. Dry it over K<sub>2</sub>CO<sub>3</sub> and distil it through a short Vigreux column. [IR: Pines et al. *J Am Chem Soc* **74** 4063 1952, Brown & Wheeler *J Am Chem Soc* **78** 2199 1956, Beilstein **1** IV 312.]

**Bromopicrin (tribromonitromethane)** [464-10-8] **M 297.8, m 10.2-10.3°, b 85-87°/16mm, d<sub>4</sub><sup>20</sup> 2.788, n<sub>D</sub><sup>20</sup> 1.579**. Steam distil it, dry it with anhydrous Na<sub>2</sub>SO<sub>4</sub> and distil it again in a vacuum. **HIGHLY TOXIC**. [Beilstein **1** H 77, **1** I 21, **1** II 43, **1** III 115, **1** IV 106.]

**R-(+)-2-Bromopropionic acid** [10009-70-8] **M 153.0, b 78°/4mm, d<sub>4</sub><sup>20</sup> 1.474, [α]<sub>D</sub><sup>25</sup> +27.2° (neat), pK<sup>25</sup> 4.07**. Dissolve it in Et<sub>2</sub>O, dry (CaCl<sub>2</sub>), evaporate and distil it through a short column. Distillation through a Podbielniak column (see S-(−)-2-chloropropionic acid below) led to decomposition. Store it in the dark under N<sub>2</sub>, preferably in sealed ampoules. Even at −10° it slowly decomposes. **LACHRYMATORY**. [Fuet et al. *J Am Chem Soc* **76** 6054 1954, Beilstein **2** IV 761.]

**3-Bromopropionic acid** [590-92-1] **M 153.0, m 62.5°, 62.5-63.5°, 63-64°, pK<sup>25</sup> 4.01**. The acid crystallises as plates from CCl<sub>4</sub>. It is soluble in organic solvents and H<sub>2</sub>O. Its *methyl ester* has **b 65°/18mm and 80°/27mm**. The *S-benzylisothiuronium salt* has **m 136°**. [Kendall & McKenzie *Org Synth Coll Vol I* 131 1941, Beilstein **2** IV 764.]

**Bromopyruvic acid (3-bromo-2-oxopropionic acid)** [1113-59-3] **M 167.0, m 79-82°, pK<sub>Est</sub> ~1.6**. Dry it by azeotropic distillation (with toluene), and then recrystallise it from dry CHCl<sub>3</sub>. Dry for 48 hours at 20° (0.5 torr) over P<sub>2</sub>O<sub>5</sub>. Store it at 0°. [Labandiniere et al. *J Org Chem* **52** 157 1987, Beilstein **3** III 1167.]

***N*-Bromosuccinimide** [128-08-5] **M 178.0, m 183-184°(dec)**. *N*-Bromosuccinimide (30g) is purified by dissolving rapidly in boiling water (300ml) and filtering through a fluted filter paper into a flask immersed in an ice bath, and left for 2 hours. The crystals are filtered off, washed thoroughly with ice-cold water (*ca* 100ml) and drained on a Büchner funnel before drying under vacuum over P<sub>2</sub>O<sub>5</sub> or CaCl<sub>2</sub> [Dauben & McCoy *J Am Chem Soc* **81** 4863 1959]. This brominating agent has also been recrystallised from acetic acid or water (10 parts see above), washed in water and dried *in vacuo* [Wilcox et al. *J Am Chem Soc* **108** 7693 1986, Shell et al. *J Am Chem Soc* **108** 121 1986, Phillips & Cohen *J Am Chem Soc* **108** 2013 1986, Beilstein **21/9** V 543.]

**Bromotetronic acid (2-bromo-4-hydroxyacetoacetic lactone)** [21151-51-9] **M 179.0, m 182.8°, pK<sup>25</sup> 2.23**. Decolourise with Norit in EtOAc, evaporate, and crystallise from EtOAc or \*C<sub>6</sub>H<sub>6</sub>. [Schuler et al. *J Phys Chem* **78** 1063 1974, Gillespie & Price *J Org Chem* **22** 782 1957, Beilstein **17** III/IV 5819.]

**Bromotrichloromethane** [75-62-7] **M 198.5, f -5.6°, m 21°, b 104.1°, d<sub>D</sub><sup>20</sup> 2.01, n<sub>D</sub><sup>20</sup> 1.5061**. Wash it with aqueous NaOH solution or dilute Na<sub>2</sub>CO<sub>3</sub>, then with H<sub>2</sub>O, and dry with CaCl<sub>2</sub>, BaO, MgSO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub> before distilling in diffuse light and storing in the dark. It has also been purified by treatment with charcoal and fractional crystallisation by partial freezing. It is purified also by vigorous stirring with portions of conc H<sub>2</sub>SO<sub>4</sub> until the acid does not discolour during several hours stirring. Wash with Na<sub>2</sub>CO<sub>3</sub> and water, dry with CaCl<sub>2</sub> and then illuminate it with a 1000W projection lamp at 15cm for 10 hours, after making it 0.01M in bromine. Pass it through a 30 × 1.5cm column of activated alumina and fractionally redistil it through a 12-in Vigreux column. [Firestone & Willard *J Am Chem Soc* **83** 3511 1961; see also Cadogan & Duell *J Chem Soc* 4154 1962, Beilstein **1** IV 77.]

**1-Bromo-2,2,2-trifluoroethane** [421-06-7] **M 163.0, m -94°, b 26-27°, d<sub>4</sub><sup>20</sup> 1.788, n<sub>D</sub><sup>20</sup> 1.332**. Wash it with water, dry (CaCl<sub>2</sub>) and distil it. [Beilstein **1** III 179, **1** IV 154.]

**Bromotrifluoromethane (Freon 13B1)** [75-63-8] **M 148.9, b -59°, d<sub>4</sub><sup>20</sup> 1.590**. Purify the gas by passing it through a tube containing P<sub>2</sub>O<sub>5</sub> on glass wool into a vacuum system where it is frozen out in a quartz tube and degassed by cycles of freezing, evacuating and thawing. [Beilstein **1** III 83, **1** IV 73.]

**5-Bromovaleric (γ-bromopentanoic) acid** [2067-33-6] **M 181.0, m 40°, pK<sub>Est</sub> ~4.6**. Crystallise the acid from petroleum ether. [Beilstein **2** IV 883.]

**(±)-Bromural [N-(aminocarbonyl)-2-bromo-3-methylbutanamide, bromisovalum]** [496-67-3] **M 223.1, m 154-155°**. Crystallise it from aqueous EtOH or toluene, and dry it in air. [Beilstein **3** H 63, **3** I 29, **3** II 51, **3** III 123, **3** IV 117.]

**1,3-Butadiene** [106-99-0] **M 54.1, b -2.6°**. Dry the gas by condensing it into a solution of triethylaluminium in decahydronaphthalene, then it is flash distilled. It has also been dried by passage over anhydrous CaCl<sub>2</sub> or distilled from NaBH<sub>4</sub>. Also purify by passing through a column packed with molecular sieves (4A), followed by cooling in a Dry-ice/MeOH bath overnight, filtering off the ice and drying over CaH<sub>2</sub> at -78° then distilling in a vacuum line. [Beilstein **1** IV 976.]

***n*-Butane** [106-97-8] **M 58.1, m -135°, b -0.5°**. Dry by passing over anhydrous Mg(ClO<sub>4</sub>)<sub>2</sub> and molecular sieves type 4A. Air is removed by prolonged and frequent degassing at -107°. [Beilstein **1** IV 236.]

**1,4-Butanediol (tetramethylene glycol)** [110-63-4] **M 90.1, f 20.4°, b 107-108°/4mm, 127°/20mm, d<sub>4</sub><sup>20</sup> 1.02, n<sub>D</sub><sup>20</sup> 1.4467**. Distil the glycol and store it over Linde type 4A molecular sieves, or crystallise it twice from anhydrous diethyl ether/acetone, and redistil it. It has been recrystallised from the melt and doubly distilled *in vacuo* in the presence of Na<sub>2</sub>SO<sub>4</sub>. [Beilstein **1** IV 2515.]

**erythro-2,3-Butanediol (meso-2,3-butylene glycol)** [5341-95-7] **M 90.1, m 32-34°, 34.4°, b 89°/16mm, 181.7°/742mm, d<sub>4</sub><sup>20</sup> 0.9939, n<sub>D</sub><sup>20</sup> 1.443, n<sub>D</sub><sup>25</sup> 1.4324**. The *meso*-form is prepared from *trans*-2,3-epoxybutane and is recrystallised from isopropyl ether at low temperature. [Wilson & Lucas *J Am Chem Soc* **58** 2396 1936, Beilstein **1** II 546, **1** III 2178, **1** IV 2524.]

**threo-2,3-Butanediol** (*rac*  $\pm$ -2,3-butylene glycol) [513-85-9] **M 90.1, m 7.6°, b 86°/16mm, 172.7°/742mm, 183-184°/760mm,  $d_4^{25}$  0.995,  $n_D^{20}$  1.443,  $n_D^{25}$  1.4310.** The  $\pm$ -*threo*-form is prepared from *cis*-2,3-epoxybutane and is recrystallised from isopropyl ether at low temperature. [Beilstein 1 II 546, 1 III 2180, 1 IV 2525.]

**threo-2,3-Butanediol** [*R,R*(-) 24347-58-8, *S,S*(+) 19132-06-0] **M 90.1, m 16-19°, 19.7°, b 77.5-78°/10mm, 179-180°/atm,  $d_4^{25}$  0.987,  $n_D^{20}$  1.443,  $n_D^{25}$  1.4310,  $[\alpha]_D^{20}$  (-) or (+) 13.1° (neat).** Purify by fractional distillation. The *bis*-(*p*-nitrobenzoate) ester has **m 141-142°, and  $[\alpha]_D^{20}$  (-) or (+) 52° (c 4, CHCl<sub>3</sub>).** [Ghirardelli & Lucas *J Am Chem Soc* 79 734 1957, Rubin et al. *J Am Chem Soc* 74 425 1952, Neish *Can J Res* 27 6 1949, Neish & Ledingham *Can J Res* 27 694 1949, Beilstein 1 IV 2524-2525.] When (-)-2*R,3R*-butane-2,3-diol (22g, 21ml, 244.5mmol) is added slowly to a stirred mixture of toluene-*p*-sulfonyl chloride (100g, 525mmol) in dry pyridine at 0° (ice-water bath), kept thus for 20 minutes, and the semi-solid mixture is set aside at ~25° overnight, then is shaken vigorously with crushed ice-water for 2 hours, poured rapidly with stirring into a mixture of concentrated HCl (70ml) and crushed ice, and the slurry is filtered gives solid (+)-2*R,3R*-butane-2,3-diol **bis(tosylate)** after washing thoroughly with H<sub>2</sub>O and drying (91.0g, 93.5%). Alternatively, the crude dry ester is dissolved in CH<sub>2</sub>Cl<sub>2</sub>, dried (MgSO<sub>4</sub>), evaporated *in vacuo*, and the residue is washed with petroleum ether (b 20-60°), and stored over solid KOH in a desiccator. It has **m 62-64° (65.1-65.5°),  $[\alpha]_D^{20}$  +37.2° (c 2.105, CHCl<sub>3</sub>).** [Corey & Mitra *J Am Chem Soc* 84 2938 1962, Lucas et al. *J Am Chem Soc* 72 2138 1950, Fryzuk & Bosnich *J Am Chem Soc* 99 6262 1977.] This (+)-2*R,3R*-ditosylate provides the chiral ligand (-)-*S,S*-CHIRAPHOS [cf 64896-28-2] after reaction with 2 mols of Ph<sub>2</sub>PLi, *via* inversion of configuration at the two chiral centres. (+)-2*R,3R*-butane-2,3-diol **bis(methanesulfonate)** has **m 123-125°,  $[\alpha]_D^{22}$  +2° (c 1, CHCl<sub>3</sub>).** [Beilstein 1 I 250, 1 II 547, 1 III 2181-3, 1 IV 2525.]

**1-Butanesulfonyl chloride** [2386-60-9] **M 156.6, b 75-76°/7mm, 98°/13mm, 100-103°/27-28mm,  $d_4^{20}$  1.2078,  $n_D^{20}$  1.4559.** It has a pungent odour and is LACHRYMATORY. If IR shows OH bands, then dissolve in Et<sub>2</sub>O, wash with cold saturated aqueous NaHCO<sub>3</sub> (care since CO<sub>2</sub> will be generated) then H<sub>2</sub>O, dry it over solid Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil the residue twice. Characterise it by shaking a solution in Et<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub> with aqueous NH<sub>3</sub>, collect the solid 1-butanesulfonamide with **m 48°** after recrystallisation from CHCl<sub>3</sub>, CCl<sub>4</sub> or Et<sub>2</sub>O/petroleum ether. [Douglass & Johnson *J Am Chem Soc* 60 1488 1938, Lee & Dougherty *J Org Chem* 5 83 1940, Beilstein 4 IV 45.]

**1-Butanethiol** [109-79-5] **M 90.2, b 98.4°,  $d_4^{25}$  0.837,  $n_D^{20}$  1.443,  $n_D^{25}$  1.440,  $pK_{Est}$  ~11.3.** Dry the thiol with CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, then reflux it over magnesium, or dry with, and distil it from CaO, under nitrogen [Roberts & Friend *J Am Chem Soc* 108 7204 1986.] It has been separated from hydrocarbons by extractive distillation with aniline.

Dissolve it also in 20% NaOH, extract with a small amount of \*C<sub>6</sub>H<sub>6</sub>, then steam distil it until clear. The distillate is then cooled and acidified slightly with 15% H<sub>2</sub>SO<sub>4</sub>. The thiol is distilled out, dried with CaSO<sub>4</sub> or CaCl<sub>2</sub>, and fractionally distilled under N<sub>2</sub> [Mathias & Filho *J Phys Chem* 62 1427 1958]. It has also been purified by precipitation as the lead mercaptide from alcoholic solution, then regeneration by addition of dilute HCl to the residue followed by steam distillation. All operations should be carried out in a fume cupboard due to the TOXICITY and obnoxious odour of the thiol. [Beilstein 1 IV 1555.]

**2-Butanethiol** [513-53-1] **M 90.2, b 37.4°/134mm,  $d_4^{25}$  0.846,  $n_D^{25}$  1.4338,  $pK_{Est}$  ~11.4.** Purify it as for 1-butanethiol. [Beilstein 1 IV 1584.]

***n*-Butanol** [71-36-3] **M 74.1, b 117.7°,  $d_4^{25}$  0.80572,  $n_D^{20}$  1.39922,  $n_D^{15}$  1.40118.** Dry it with MgSO<sub>4</sub>, CaO, K<sub>2</sub>CO<sub>3</sub>, or solid NaOH, followed by refluxing with, and distilling from, small amounts of calcium, magnesium activated with iodine, or aluminium amalgam. It can also be dried with molecular sieves, or by refluxing with *n*-butyl phthalate or succinate. (For method, see *Ethanol*.) *n*-Butanol can also be dried by efficient fractional distillation, water passing over in the first fraction as a binary azeotrope (contains about 37% water). An ultraviolet-transparent distillate has been obtained by drying with magnesium and distilling from sulfanilic acid. To remove bases, aldehydes and ketones, the alcohol is washed with dilute H<sub>2</sub>SO<sub>4</sub>, then NaHSO<sub>4</sub> solution; esters are removed by boiling for 1.5 hours with 10% NaOH.

It has also been purified by adding NaBH<sub>4</sub> (2g) to butanol, (1.5L) gently bubbling under argon and a reflux con-

denser for 1 day at 50°. Then add freshly cut sodium (2g, washed with butanol) and reflux 1 day. Distil, and collect the middle fraction [Jou & Freeman *J Phys Chem* **81** 909 1977]. [Beilstein 1 IV 1506.]

**2-Butanone (methyl ethyl ketone, MEK)** [78-93-0] **M 72.1, b 79.6°, d<sub>4</sub><sup>20</sup> 0.853, n<sub>D</sub><sup>20</sup> 1.37850, n<sub>D</sub><sup>25</sup> 1.37612, pK<sup>25</sup> -7.2 (aqueous H<sub>2</sub>SO<sub>4</sub>).** In general, purification methods are the same as for acetone. Aldehydes can be removed by refluxing with KMnO<sub>4</sub> + CaO, until the Schiff aldehyde test is negative, prior to distillation. Shaking with saturated K<sub>2</sub>CO<sub>3</sub>, or passing through a small column of activated alumina, removes cyclic impurities. The ketone can be dried by careful distillation (an azeotrope containing 11% water boils at 73.4°), or over CaSO<sub>4</sub>, P<sub>2</sub>O<sub>5</sub>, Na<sub>2</sub>SO<sub>4</sub>, or K<sub>2</sub>CO<sub>3</sub>, followed by fractional distillation. Purification as the bisulfite addition compound is achieved by shaking with excess saturated Na<sub>2</sub>SO<sub>3</sub>, cooled to 0°, filtering off the precipitate, washing with a little ethyl ether and drying in air; this is followed by decomposition with a slight excess of Na<sub>2</sub>CO<sub>3</sub> solution and steam distillation, the distillate being saturated with K<sub>2</sub>CO<sub>3</sub> so that the ketone can be separated, dried with K<sub>2</sub>CO<sub>3</sub>, filtered, and distilled. Purification as the *Nal addition compound* (**m 73-74°**) is more convenient. (For details, see *Acetone*.) Small quantities of 2-butanone can be purified by conversion to the semicarbazone, recrystallisation to constant melting point, drying under vacuum over CaCl<sub>2</sub> and paraffin wax, refluxing for 30 minutes with excess oxalic acid, followed by steam distillation, salting out, drying and distilling. [Cowan et al. *J Chem Soc* 171 1940]. [Beilstein 1 IV 3243.]

**cis-2-Butene** [590-18-1] **M 56.1, b 2.95-3.05°/746mm.** The gas is dried with CaH<sub>2</sub> and purified by gas chromatography. [Beilstein 1 H 205, 1 II 176, 1 III 728, 1 IV 778.] **HIGHLY FLAMMABLE.**

**trans-2-Butene** [624-64-6] **M 56.1, b 0.3-0.4°/744mm.** The gas is dried with CaH<sub>2</sub> and purified by gas chromatography. [Beilstein 1 H 205, 1 II 176, 1 III 730, 1 IV 781.] **HIGHLY FLAMMABLE.**

**2-Butene-1,4-dicarboxylic acid (trans-3-hexenedioic acid, trans-β-hydromuconic acid)** [4436-74-2] **M 144.1, m 194-197°, 195-196°, pK<sub>Est(1)</sub> ~4.2, pK<sub>Est(2)</sub> ~5.00.** Crystallise the acid from boiling water, then dry it at 50-60° in a vacuum oven. [Beilstein 2 IV 2237.]

**2-Butoxyethanol (butyl cellosolve)** [111-76-2] **M 118.2, b 171°/745mm, d<sub>4</sub><sup>20</sup> 0.903, n<sub>D</sub><sup>20</sup> 1.4191.** Peroxides can be removed by refluxing with anhydrous SnCl<sub>2</sub> or by passage under slight pressure through a column of activated alumina. Dry with anhydrous K<sub>2</sub>CO<sub>3</sub> and CaSO<sub>4</sub>, filter and distil, or reflux with, and distil from NaOH. [Beilstein 1 IV 2380.]

**n-Butyl acetate** [123-86-4] **M 116.2, b 126.1°, d<sub>4</sub><sup>20</sup> 0.882, n<sub>D</sub><sup>20</sup> 1.394.** Distil, reflux with successive portions of KMnO<sub>4</sub> until the colour persists, dry with anhydrous CaSO<sub>4</sub>, filter and redistil. [Beilstein 2 IV 143.]

**tert-Butyl acetate** [540-88-5] **M 116.2, b 97-98°, d<sub>4</sub><sup>20</sup> 0.866, n<sub>D</sub><sup>20</sup> 1.387.** Wash the ester with 5% Na<sub>2</sub>CO<sub>3</sub> solution, then saturated aqueous CaCl<sub>2</sub>, dry with CaSO<sub>4</sub> and distil it. [McClosky et al. *Org Synth Coll Vol IV* 263 1963, Mangia et al. *Org Prep Proc Int* **18** 13 1986, Beilstein 2 IV 151.]

**tert-Butyl acetoacetate** [1694-31-1] **M 158.2, b 71°/10mm, 85°/20mm, d<sub>4</sub><sup>20</sup> 0.954, n<sub>D</sub><sup>20</sup> 1.42.** Distil it under reduced pressure through a short column. [Lawesson et al. *Org Synth Coll Vol V* 155 1973, Lawesson et al. *Org Synth* **42** 28 1962, Beilstein 3 IV 1536.] **HARMFUL VAPOUR.**

**tert-Butylacetyl chloride** [7065-46-5] **M 134.6, b 68-71°/100mm, 81°/180mm, 128-132°/atm, d<sub>4</sub><sup>20</sup> 0.964, n<sub>D</sub><sup>20</sup> 1.423.** Distil it under vacuum. If IR shows OH group, then treat with thionyl chloride or oxalyl chloride at ca 50° for 30 minutes, evaporate and fractionate the residue using a short column. Strongly **LACHRYMATORY**, use a good fume hood. [Berliner & Berliner *J Am Chem Soc* **72** 222 1950, Traynham & Battiste *J Org Chem* **22** 1551 1957, Beilstein 2 IV 956.]

**Butyl acrylate** [141-32-2] **M 128.2, b 59°/25mm, d<sub>4</sub><sup>20</sup> 0.894, n<sub>D</sub><sup>12</sup> 1.4254.** Wash it repeatedly with aqueous NaOH to remove inhibitors such as hydroquinone, then with distilled water. Dry with CaCl<sub>2</sub>. Fractionally distil under reduced pressure in an all-glass apparatus. The middle fraction is sealed under N<sub>2</sub> and stored at 0° in the dark until required or with a stabiliser [Mallik & Das *J Am Chem Soc* **82** 4269 1960]. [Beilstein 2 IV 1463.]

***tert*-Butyl acrylate** [1663-39-4] **M 128.2, b 30.0-30.8°/26mm, 61-63°/15mm, 117-120°/760mm,  $d_4^{25}$  0.875,  $n_D^{20}$  1.410.** Purify the ester by fractional distillation. If it contains acid (OH bands in the IR), then dissolve it in Et<sub>2</sub>O, wash it with aqueous NaHCO<sub>3</sub>, dry the organic layer (Na<sub>2</sub>SO<sub>4</sub>), filter it and distil it under reduced pressure. Stabilise it by adding hydroquinone monomethyl ether (~0.05%). It forms a crystalline *tert*-butyl acrylate polymer which is soluble in organic solvents [Garrett et al. *J Am Chem Soc* **81** 1007 1959]. [Beilstein **2** IV 1465.] For other alkyl acrylates see Rehberg *Org Synth Coll Vol* **3** 146 1955.

**(±)-*sec*-Butyl alcohol (± 2-butanol)** [78-92-2; 15892-23-6] **M 74.1, b 99.4°,  $d_4^{20}$  0.808.** Purification methods are the same as for *n*-Butanol. These include drying with K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, followed by filtration and fractional distillation, refluxing with CaO, distillation, then refluxing with magnesium and redistillation, and refluxing with, then distilling from CaH<sub>2</sub>. Calcium carbide has also been used as a drying agent. The anhydrous alcohol is obtained by refluxing with *sec*-butyl phthalate or succinate. (For method see *Ethanol*.) Small amounts of alcohol can be purified *via* conversion to the alkyl hydrogen phthalate and recrystallisation [Hargreaves *J Chem Soc* 3679 1956]. For purification of optical isomers, see Timmermans and Martin [*J Chem Phys* **25** 411 1928]. [Beilstein **2** III 1566.]

***tert*-Butyl alcohol** [75-65-0] **M 74.1, m 23-25°, 25.7°, b 28.3°/60mm, 43.3°/123.8mm, 61.8°/315mm, 72.5°/507mm, 82.45°/760mm,  $d_4^{20}$  0.7858,  $n_D^{20}$  1.3878.** It is synthesised commercially by the hydration of 2-methylpropene in dilute H<sub>2</sub>SO<sub>4</sub>. Dry it with CaO, K<sub>2</sub>CO<sub>3</sub>, CaSO<sub>4</sub> or MgSO<sub>4</sub>, filter and fractionally distil it. Dry further by refluxing with, and distilling from, either magnesium activated with iodine, or small amounts of calcium, sodium or potassium, under nitrogen. Passage through a column of type 4A molecular sieve is another effective method of drying; as well as refluxing with *tert*-butyl phthalate or succinate. (For method see *Ethanol*.) Other methods include refluxing with excess aluminium *tert*-butylate, or standing with CaH<sub>2</sub>, and distilling as needed. Further purification is achieved by fractional crystallisation by partial freezing, taking care to exclude moisture. *tert*-Butyl alcohol samples containing much water can be dried by adding \*benzene, so that the water distils off as a tertiary azeotrope, **b 67.3°**. Traces of isobutylene have been removed from dry *tert*-butyl alcohol by bubbling dry pre-purified nitrogen through for several hours at 40-50° before using. It forms azeotropic mixtures with a large number of compounds. It has also been purified by distillation from CaH<sub>2</sub> into Linde 4A molecular sieves which had been activated at 350° for 24 hours [Jaeger et al. *J Am Chem Soc* **101** 717 1979]. [Beilstein **1** IV 1609.]

**Rapid purification:** Dry *tert*-butanol over CaH<sub>2</sub> (5% w/v), distil and store it over 3A molecular sieves.

***n*-Butylamine** [109-73-9] **M 73.1, b 77.8°,  $d_4^{20}$  0.740,  $n_D^{20}$  1.4009,  $n_D^{25}$  1.399,  $pK^{25}$  10.66.** Dry it with solid KOH, K<sub>2</sub>CO<sub>3</sub>, LiAlH<sub>4</sub>, CaH<sub>2</sub> or MgSO<sub>4</sub>, then reflux it with, and fractionally distil it from P<sub>2</sub>O<sub>5</sub>, CaH<sub>2</sub>, CaO or BaO. Further purification is by precipitation as the *hydrochloride*, **m 213-213.5°**, from ethereal solution by bubbling HCl gas into it. This is re-precipitated three times from EtOH by adding ether, followed by liberation of the free amine using excess strong base. The amine is extracted into ether, which is separated, dried with solid KOH, the ether is removed by evaporation and then the amine is distilled. It is stored in a desiccator over solid NaOH [Bunnett & Davis *J Am Chem Soc* **82** 665 1960, Lycan et al. *Org Synth Coll Vol* **II** 319 1943]. [Beilstein **4** IV 540.] **SKIN IRRITANT.**

***R*-(-)-*sec*-Butylamine** [13250-12-9] **M 73.1, b 61-63°/atm, 62.5°/atm,  $d_4^{20}$  0.731,  $n_D^{20}$  1.393,  $[\alpha]_D^{20}$  -7.5° (neat),  $pK^{25}$  10.56.** Dry it over solid NaOH overnight and fractionate it through a short helices packed column. The *L*-hydrogen tartrate salt has **m 139-140°** (from H<sub>2</sub>O), the *1H<sub>2</sub>O* has **m 96°**  $[\alpha]_D^{21}$  +18.1° (c 11, H<sub>2</sub>O), the *hydrochloride* has **m 152°**  $[\alpha]_D^{21}$  -1.1° (c 13, H<sub>2</sub>O) and the *benzoyl* derivative crystallises from EtOH as needles\* with **m 97°**,  $[\alpha]_D^{21}$  -34.9° (c 11, H<sub>2</sub>O). [Bruck et al. *J Chem Soc* 921 1956, Kjaer & Hansen *Acta Chem Scand* **11** 898 1957.] [Beilstein **4** H 161, **4** I 372, **4** III 308, **4** IV 617.] The *S*-(+)-**enantiomer** has same properties except for the optical rotation which has the opposite sign.

***tert*-Butylamine** [75-64-9] **M 73.1, b 42°,  $d_4^{20}$  0.696,  $pK^{25}$  10.68.** Dry it with KOH or LiAlH<sub>4</sub>, and/or distil it from CaH<sub>2</sub> or BaO. [Beilstein **4** IV 657.]

***tert*-Butylammonium bromide** [60469-70-7] **M 154.1, m >250°(dec).** Recrystallise the salt several times from absolute EtOH and thoroughly dry it at 105° *in vacuo*. [Beilstein **4** IV 657.]

***n*-Butyl bromide** [109-65-9] M 137.0, b 101-102°, d<sup>25</sup> 1.2678, n<sub>D</sub><sup>20</sup> 1.4399, n<sup>25</sup> 1.4374. Wash the bromide with conc H<sub>2</sub>SO<sub>4</sub>, water, 10% Na<sub>2</sub>CO<sub>3</sub> and again with water. Dry it over CaCl<sub>2</sub>, CaSO<sub>4</sub> or K<sub>2</sub>CO<sub>3</sub>, and distil it. Redistil it after drying with P<sub>2</sub>O<sub>5</sub>, or pass it through two columns containing 5:1 silica gel/Celite mixture and store it with freshly activated alumina. [Beilstein 1 IV 258.]

***tert*-Butyl bromoacetate** [5292-43-3] M 195.1, b 52°/10mm, 74-76°/25mm, d<sub>4</sub><sup>20</sup> 1.324, n<sub>D</sub><sup>25</sup> 1.4162. Dissolve the ester in Et<sub>2</sub>O, wash it well with ice cold 10% aqueous K<sub>2</sub>CO<sub>3</sub>, dry it over CaCl<sub>2</sub>, filter and evaporate the Et<sub>2</sub>O, then fractionate it through a Vigreux column in a vacuum. LACHRYMATORY. [Abramovitch et al. *J Am Chem Soc* 64 2274 1942, Abramovitch & Hauser *J Am Chem Soc* 65 986 1943, Beilstein 2 III 482.]

***tert*-Butyl carbazate** [870-46-2] M 132.2, m 41-42°, b 64°/0.01mm, 55-57°/0.4mm. Distil it in a Claisen flask with a water or oil bath at ca 80°. After a couple of drops have distilled, the carbazate is collected as an oil which solidifies to a snow white solid. It can be crystallised with 90% recovery from a 1:1 mixture of petroleum ether (b 30-60°) and petroleum ether (b 60-70°). [Carpino et al. *Org Synth Coll Vol V* 166 1973, Caprino et al. *Org Synth* 44 20 1964, Beilstein 3 IV 175.]

***n*-Butyl chloride** [109-69-3] M 92.6, b 78°, d<sub>4</sub><sup>20</sup> 0.886, n<sub>D</sub><sup>25</sup> 1.4021. Shake it repeatedly with conc H<sub>2</sub>SO<sub>4</sub> (until no further colour develops in the acid layer), then wash it with water, aqueous NaHCO<sub>3</sub> or Na<sub>2</sub>CO<sub>3</sub>, and more water. Dry it with CaCl<sub>2</sub>, or MgSO<sub>4</sub> (then with P<sub>2</sub>O<sub>5</sub> if desired), decant and fractionally distil it. Alternatively, a stream of oxygen continuing ca three times as long as is necessary to obtain the first coloration of starch iodide paper by the exit gas. After washing with NaHCO<sub>3</sub> solution to hydrolyse ozonides and to remove the resulting organic acid, the liquid is dried and distilled [Chien & Willard *J Am Chem Soc* 75 6160 1953]. [Beilstein 1 IV 246.]

***tert*-Butyl chloride** [507-20-0] M 92.6, f -24.6°, b 50.4°, d<sub>4</sub><sup>20</sup> 0.851, n<sub>D</sub><sup>25</sup> 1.38564. Purification methods commonly used for other alkyl halides lead to decomposition. Some impurities can be removed by photochlorination with a small amount of chlorine prior to use. The liquid is washed with ice water, dried with CaCl<sub>2</sub> or CaCl<sub>2</sub> + CaO and fractionally distilled. It has been further purified by repeated fractional crystallisation by partial freezing. [Beilstein 1 IV 288.]

***tert*-Butyl chloroacetate** [107-59-5] M 150.6, b 48-49°/11mm, 60.2°/15mm, 155°/atm (dec), d<sub>4</sub><sup>25</sup> 1.4204, n<sub>D</sub><sup>20</sup> 1.4259. Check the NMR spectrum; if satisfactory then distil in a vacuum; if not then dissolve in Et<sub>2</sub>O, wash with H<sub>2</sub>O, 10% H<sub>2</sub>SO<sub>4</sub> until the acid extract does not become cloudy when made alkaline with NaOH. Wash the organic layer again with H<sub>2</sub>O, then saturated aqueous NaHCO<sub>3</sub>, dry over Na<sub>2</sub>SO<sub>4</sub>, evaporate and fractionate it through a carborundum-packed column or a 6-inch Widmer column (see *tert-butyl ethyl malonate for precautions to avoid decomposition during distillation*). [Johnson et al. *J Am Chem Soc* 75 4995 1953, Baker *Org Synth Coll Vol III* 144 1944, Beilstein 2 III 444.]

***tert*-Butyl cyanide (trimethylacetonitrile, pivalonitrile)** [630-18-2] M 83.1, m 16-18°, d<sub>4</sub><sup>20</sup> 0.765, b 104-106°. Purify it by a two-stage vacuum distillation and de-gas by the freeze-pump-thaw technique. Store it under vacuum at 0°. TOXIC, use an efficient fume hood. [Beilstein 2 IV 875.]

***tert*-Butyl cyanoacetate** [1116-98-9] M 141.2, b 40-42°/0.1mm, 54-56°/0.3mm, 90°/10mm, 107-108°/23mm, d<sub>4</sub><sup>20</sup> 0.989, n<sub>D</sub><sup>20</sup> 1.4198. The IR spectrum of a film should have bands at 1742 (ester CO) and 2273 (C≡N), but no band at ca 3500 broad (OH) cm<sup>-1</sup>. If it does not have the last-named band, then fractionally distil; otherwise dissolve in Et<sub>2</sub>O, wash with saturated aqueous NaHCO<sub>3</sub>, dry over K<sub>2</sub>CO<sub>3</sub>, evaporate Et<sub>2</sub>O, and distil the residue under a vacuum (see *tert-butyl ethyl malonate for precautions to avoid decomposition during distillation*). [Beech & Piggott *J Chem Soc* 423 1955, Dahn & Hauth *Helv Chim Acta* 42 1214 1959, Beilstein 2 I 255.]

***tert*-Butyl diazoacetate** [3505950-8] M 142.2, b 51-53°/12mm, d<sub>4</sub><sup>25</sup> 1.026, n<sub>D</sub><sup>20</sup> 1.443. It is a poisonous orange-yellow liquid which is explosive, and the necessary precautions should be strictly adhered to (i.e. efficient fume cupboard, and face and body protection; see reference below). Check the purity by TLC on Merck Kieselgel F<sub>254</sub> or Eastman Kodak Silica Gel without indicator using CHCl<sub>3</sub> as eluent (R<sub>F</sub> 0.72). If the

ester is suspect, then dissolve it in Et<sub>2</sub>O, wash it with brine, H<sub>2</sub>O, dry the organic layer (Na<sub>2</sub>SO<sub>4</sub>), and filter. Remove most of the Et<sub>2</sub>O at 30°/water pump vacuum (rotovap), the remaining ether during distillation under lower pressure (<12mm), and the residual oil then distils as a coloured liquid using a water bath at ~50° (no higher as heat source). [Regitz et al. *Org Synth Coll Vol V* 179 1972.]

***n*-Butyl disulfide** [629-45-8] M 178.4, b 110-113°/15mm, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>22</sup> 1.494. Shake it with lead peroxide, filter and distil it in a vacuum under N<sub>2</sub>. [Beilstein 1 IV 1560.]

***n*-Butyl ether (di-*n*-butyl ether)** [142-96-1] M 130.2, b 52-53°/26mm, 142.0°/760mm, d<sub>4</sub><sup>20</sup> 0.764, n<sub>D</sub><sup>20</sup> 1.39925, n<sub>D</sub><sup>25</sup> 1.39685, pK<sup>25</sup> -5.40 (aqueous H<sub>2</sub>SO<sub>4</sub>). Peroxides (detected by the liberation of iodine from weakly acid HCl solutions of 2% KI) can be removed by shaking 1L of ether with 5-10ml of a solution comprising of ferrous sulfate (6.0g) in conc H<sub>2</sub>SO<sub>4</sub> (6ml) of water (110ml), with aqueous Na<sub>2</sub>SO<sub>3</sub>, or with acidified NaI, water, then aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>. After washing with dilute NaOH, KOH, or Na<sub>2</sub>CO<sub>3</sub>, then water, the ether is dried with CaCl<sub>2</sub> and distilled. It can be further dried by distillation from CaH<sub>2</sub> or Na (after drying with P<sub>2</sub>O<sub>5</sub>), and stored in the dark with Na or NaH. The ether can also be purified by treating with CS<sub>2</sub> and NaOH, expelling the excess sulfide by heating. The ether is then washed with water, dried with NaOH and distilled [Kusama & Koike *J Chem Soc Jpn, Pure Chem Sect* 72 229 1951]. Other purification procedures include passage through an activated alumina column to remove peroxides, or through a column of silica gel, and distillation after adding about 3% (v/v) of a 1M solution of MeMgI in *n*-butyl ether. [Beilstein 1 IV 1520.]

***n*-Butyl ethyl ether** [628-81-9] M 102.2, b 92.7°, d<sub>4</sub><sup>20</sup> 0.751, n<sub>D</sub><sup>20</sup> 1.38175, n<sub>D</sub><sup>25</sup> 1.3800. Purify by drying with CaSO<sub>4</sub>, by passage through a column of activated alumina (to remove peroxides), followed by prolonged refluxing with Na and then fractional distillation. [Beilstein 4 IV 1518.]

***tert*-Butyl ethyl ether** [637-92-3] M 102.2, b 71-72°, d<sub>4</sub><sup>20</sup> 0.741. Dry the ether with CaSO<sub>4</sub>, pass it through an alumina column, and fractionally distil it. [Beilstein 1 IV 1618.]

***tert*-Butyl ethyl malonate** [32864-38-3] M 188.2, b 83-85°/8mm, 93-95°/17mm, 107-109°/24mm, d<sub>4</sub><sup>25</sup> 0.994, n<sub>D</sub><sup>24</sup> 1.4150. A likely impurity is monoethyl malonate; check IR for OH bands at 3330 br cm<sup>-1</sup>. To ca 50g of ester add ice cold NaOH (50g in 200ml of H<sub>2</sub>O and 200g of ice). Swirl a few times (filter off ice if necessary), place it in a separating funnel and extract with Et<sub>2</sub>O (2 x 75ml). Dry the extract (MgSO<sub>4</sub>) (since traces of acid decompose the *t*-Bu group of the ester, the distillation flask has to be washed with aqueous NaOH, rinsed with H<sub>2</sub>O and allowed to dry). Addition of some K<sub>2</sub>CO<sub>3</sub> or MgO before distilling is recommended to inhibit decomposition. Distil it under reduced pressure through a 10cm Vigreux column. *Decomposition is evidenced by severe foaming due to autocatalytic decomposition and cannot be prevented from accelerating except by stopping the distillation and rewashing the distillation flask with alkali again.* [Breslow et al. *J Am Chem Soc* 66 1287 1944, Hauser et al. *J Am Chem Soc* 64 2714 1942, Strube *Org Synth Coll Vol IV* 417 1963, Stube *Org Synth* 37 35 1957, Beilstein 2 IV 1884.]

***n*-Butyl formate** [592-84-7] M 102.1, b 106.6°, d<sub>4</sub><sup>20</sup> 0.891, n<sub>D</sub><sup>20</sup> 1.3890. Wash the formate with saturated NaHCO<sub>3</sub> solution in the presence of saturated NaCl, until no further reaction occurs, then with saturated NaCl solution, dry (MgSO<sub>4</sub>), filter and fractionally distil the filtrate. [Beilstein 2 IV 28.]

**Butyl glycolate** [7397-62-8] M 132.2, b 191-192°/755mm, 187-190°/atm, d<sub>4</sub><sup>20</sup> 1.019, n<sub>D</sub><sup>20</sup> 1.4263. Dissolve the ester in CHCl<sub>3</sub> (EtOH-free), wash with 5% KHCO<sub>3</sub> until effervescence ceases (if free acid is present), dry over CaCl<sub>2</sub>, filter, evaporate and distil through a short column. [Böhme & Opfer *Z Anal Chem* 139 255 1953, Filachione et al. *J Am Chem Soc* 73 5265 1951, Beilstein 3 IV 589.]

***tert*-Butyl hydroperoxide (TBHP)** [75-91-2] M 90.1, f 5.4°, m 0.5-2.0°, b 38°/18mm, d<sub>4</sub><sup>20</sup> 0.900, n<sub>D</sub><sup>20</sup> 1.4013, pK<sup>20</sup> 12.8. *Care should be taken when handling this peroxide because of the possibility of EXPLOSION. It explodes when heated over an open flame. Work in an efficient fume cupboard, behind a thick plastic transparent shield, and with eye protection.* Alcoholic and volatile impurities can be removed by prolonged refluxing at 40° under reduced pressure, or by steam distillation. For example, Bartlett, Benzing and Pincock [*J Am Chem Soc* 82 1762 1960] refluxed at 30mm pressure in an apparatus for azeotropic liquid separation until the

two phases no longer separated, and then distilled at 41°/23mm. Pure material is stored under N<sub>2</sub>, in the dark at 0°. Crude commercial material has been added to 25% NaOH below 30°, and the crystals of the sodium salt have been collected, washed twice with \*benzene and dissolved in distilled water. After the pH of the solution is adjusted to 7.5 by adding solid CO<sub>2</sub>, the peroxide is extracted into petroleum ether, from which, after drying with K<sub>2</sub>CO<sub>3</sub>, TBHP is recovered by distilling off the solvent under reduced pressure at room temperature [O'Brien et al. *J Am Chem Soc* **79** 6238 1957]. **The temperatures should be kept below 75°.** It has also been distilled through a helices packed column (*ca* 15 plates) and the material with **b 34-35°/20mm** is collected. Similarly, a solution in petroleum ether has been extracted with cold aqueous NaOH, and the hydroperoxide has been regenerated by adding at 0°, KHSO<sub>4</sub> to a pH not higher than 4.5, then extracted into diethyl ether, dried with MgSO<sub>4</sub>, filtered and the ether evaporated in a rotary evaporator under reduced pressure at as low a temperature as possible [Milac & Djokic *J Am Chem Soc* **84** 3098 1962].

A 3M solution of TBHP in CH<sub>2</sub>Cl<sub>2</sub> is prepared by swirling of commercial TBHP (85ml, 0.61mol, 70% TBHP-30% H<sub>2</sub>O, **d** 0.935 *ca* 7.2mmol/ml) with of CH<sub>2</sub>Cl<sub>2</sub> (140ml) in a separating funnel. The milky mixture is allowed to stand until the phases separate (*ca* 30 minutes). The organic (lower) layer (*ca* 200ml) containing 0.60mole of TBHP is separated from the aqueous layer (*ca* 21ml) and used without further drying. TBHP is assayed by iodometric titration. With 90% grade TBHP (w/w, **d** 0.90, *ca* 9.0mmole/ml) no separation of layers occurs, i.e. when TBHP (66.67ml, 0.60mole) is added to CH<sub>2</sub>Cl<sub>2</sub> (140ml) the resulting solution (*ca* 200ml) should be clear. [Walling & Buckler *J Am Chem Soc* **77** 6032 1955, Rogers & Campbell *J Am Chem Soc* **74** 4742 1952, Akashi et al. *J Org Chem* **43** 2063 1978 state the quality of available grades, handling and compatibility for reactions, *Beilstein* **1** IV 1616.]

**n-Butyl iodide (1-iodobutane)** [542-69-8] **M 184.0, b 130.4°, d**<sub>4</sub><sup>20</sup> **1.616, n**<sub>D</sub><sup>25</sup> **1.44967.** Dry the iodide with MgSO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub>, fractionally distil it through a column packed with glass helices, taking the middle fraction and storing over calcium or mercury in the dark. *Alternatively*, purify it by prior passage through activated alumina or by shaking with conc H<sub>2</sub>SO<sub>4</sub> then washing with Na<sub>2</sub>SO<sub>3</sub> solution. It has also been treated carefully with sodium to remove free HI and H<sub>2</sub>O, before distilling through a column containing copper turnings at the top. Another purification procedure consisted of treatment with bromine, followed by extraction of free halogen with Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, washing with H<sub>2</sub>O, drying and fractionally distilling. [*Beilstein* **1** IV 271.]

**tert-Butyl iodide** [558-17-8] **M 184.0, b 100°(dec), d**<sub>4</sub><sup>20</sup> **1.544.** Vacuum distillation has been used to obtain a distillate which remained colourless for several weeks at -5°. More extensive treatment has been used by Boggs, Thompson and Crain [*J Phys Chem* **61** 625 1957] who washed it with aqueous NaHSO<sub>3</sub> solution to remove free iodine, dried this for 1 hour over Na<sub>2</sub>SO<sub>3</sub> at 0°, and purified it by four or five successive partial freezings of the liquid to obtain colourless material, and was stored at -78° with Ag wool. [*Beilstein* **1** IV 300.]

**tert-Butyl isocyanate** [1609-86-5] **M 99.1, m 10.5-11.5°, b 30.5-32°/10mm, 64°/52mm, d**<sub>25</sub><sup>25</sup> **0.9079, n**<sub>D</sub><sup>25</sup> **1.470.** It is **LACHRYMATORY** and **TOXIC**, and should have IR with 2251 (C≡N) cm<sup>-1</sup> but no OH bands. The NMR should have one band at 1.37 ppm from TMS. Purify it by fractional distillation under reduced pressure. [Greene & Bergmark *J Org Chem* **36** 3056 1971, Curtius *J Prakt Chem* **125** 152 1930, *Beilstein* **4** IV 669.]

**tert-Butyl isocyanide** [7188-38-7] **M 83.1, b 91-92°/730mm, 90°/758mm, d**<sub>4</sub><sup>20</sup> **0.735.** Dissolve it in petroleum ether (b 40-60°), wash it with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, remove petroleum ether under slight vacuum, and distil it using a vacuum-jacketed Vigreux column at atmospheric pressure, its IR has a band at 2134 cm<sup>-1</sup>. [Ugi & Meyr *Chem Ber* **93** 239 1960, *Beilstein* **4** IV 661.] It has **toxic vapours**.

**tert-butyl isocyanoacetate** [2769-72-4] **M 141.2, b 50°/0.1mm, 49-50°/10mm, 63-65°/15mm, d**<sub>4</sub><sup>20</sup> **0.970, n**<sub>D</sub><sup>20</sup> **1.420.** If it contains some free acid (OH bands in IR), then dissolve it in Et<sub>2</sub>O, shake with 20% Na<sub>2</sub>CO<sub>3</sub>, dry over anhydrous K<sub>2</sub>CO<sub>3</sub>, evaporate and distil it. [Ugi et al. *Chem Ber* **94** 2814 1961, Schöllkopf *Angew Chem* **89** 351 1977.]

**n-Butyl methacrylate** [97-88-1] **M 142.2, b 49-52°/0.1mm, 163-165°/atm, d**<sub>4</sub><sup>20</sup> **0.896, n**<sub>D</sub><sup>20</sup> **1.424.** Purify it as for butyl acrylate. [*Beilstein* **2** IV 1525.]

**tert-Butyl methacrylate** [585-07-9] M 142.2, f  $-48^{\circ}$ , b  $135-136^{\circ}/760\text{mm}$ ,  $d_4^{20}$  0.878,  $n_D^{20}$  1.415. Purify it as for butyl acrylate. [Beilstein 2 IV 1582.]

**n-Butyl methyl ether** [628-28-4] M 88.2, b  $70^{\circ}$ ,  $d_4^{20}$  0.744,  $pK^{25}$   $-3.50$  (aqueous  $H_2SO_4$ ). Dry it with  $CaSO_4$ , pass it through an alumina column to remove peroxides, and fractionally distil it. [Beilstein 1 IV 1518.]

**tert-Butyl methyl ether** (methyl *tert*-butyl ether, MTBE) [1634-04-4] M 88.2, b  $56^{\circ}$ , n 1.369. Purify it as for *n*-butyl methyl ether. [Beilstein 1 IV 1615.]

**tert-Butyl methyl ketone** (3,3-dimethyl-2-butanone, pinacolone) [75-97-8] M 100.2, b  $105^{\circ}/746\text{mm}$ ,  $106^{\circ}/760\text{mm}$ ,  $d_4^{20}$  0.814,  $n_D^{20}$  1.401. Reflux the ketone with a little  $KMnO_4$ . Dry it with  $CaSO_4$  and distil it. [Beilstein 1 IV 3310.]

**tert-Butyl nitrite** [540-80-7] M 103.1, b  $34^{\circ}/250\text{mm}$ ,  $61-63^{\circ}/\text{atm}$ ,  $d_4^{20}$  0.8671,  $n_D^{25}$  1.3660. If it is free from OH bands (IR) then distil it through a 12inch helices packed column under reduced pressure, otherwise wash with aqueous 5%  $NaHCO_3$  (effervescence), then  $H_2O$ , dry ( $Na_2SO_4$ ) and fractionate it through a 10 theoretical plates column at *ca* 10mm pressure. [Allen *J Chem Soc* 1968 1954, Coe & Doumani *J Am Chem Soc* 70 1516 1948, UV: Ungnade & Smiley *J Org Chem* 21 993 1956, IR: Terte *Bull Soc Chim Belg* 60 240 1951, Beilstein 1 IV 1622.]

**tert-Butyl peracetate** [107-71-1] M 132.2, b  $23-24^{\circ}/0.5\text{mm}$ ,  $n_D^{25}$  1.4030. Wash the ester with  $NaHCO_3$  from a \*benzene solution, then redistil to remove \*benzene [Kochi *J Am Chem Soc* 84 774 1962]. Handle with adequate protection due to possible **EXPLOSIVE** nature. [Beilstein 2 IV 391.]

**tert-Butylperoxy isobutyrate** [109-13-7] M 160.2, f  $-45.6^{\circ}$ . After diluting the material (90ml) with petroleum ether (120ml), the mixture is cooled to  $5^{\circ}$  and shaken twice with 5%  $NaOH$  solution (90ml portions, also at  $5^{\circ}$ ). The non-aqueous layer, after washing once with cold water, is dried at  $0^{\circ}$  with a mixture of anhydrous  $MgSO_4$  and  $MgCO_3$  containing *ca* 40%  $MgO$ . After filtering, this material is passed, twice, through a column of silica gel at  $0^{\circ}$  (to remove *tert*-butyl hydroperoxide). The solution is then evaporated at  $0^{\circ}/0.5-1\text{mm}$  to remove the solvent, and the residue is recrystallised several times from petroleum ether at  $-60^{\circ}$ , then subjected to high vacuum to remove traces of solvent [Milos & Golubovic *J Am Chem Soc* 80 5994 1958]. Handle with adequate protection due to possible **EXPLOSIVE** nature.

**Butyl stearate** [123-95-5] M 340.6, m  $26.3^{\circ}$ ,  $d_4^{20}$  0.861. Acidic impurities are removed by shaking with 0.05M  $NaOH$  or a 2%  $NaHCO_3$  solution, followed by several water washes, then purified by fractional freezing of the melt and fractional crystallisation from solvents with boiling points below  $100^{\circ}$ . [Beilstein 2 IV 1219.]

**S-tert-Butyl thioacetate** [999-90-6] M 132.2, b  $31-32^{\circ}/11\text{mm}$ ,  $38^{\circ}/14\text{mm}$ ,  $44-45^{\circ}/28\text{mm}$ ,  $67^{\circ}/54\text{mm}$ ,  $135.6-135.9^{\circ}/773\text{mm}$ ,  $d_4^{25}$  0.9207,  $n_D^{20}$  1.4532. Dissolve it in  $CHCl_3$  (EtOH-free), wash with  $H_2O$ , 10%  $H_2SO_4$ , saturated aqueous  $NaHCO_3$  (care  $CO_2$  liberated),  $H_2O$  again, dry over Drierite and anhydrous  $K_2CO_3$ , and fractionate under reduced pressure. [Rylander & Tarbell *J Am Chem Soc* 72 3021 1950, Beilstein 2 IV 546.]

**N-tert-Butyl urea** [1118-12-3] M 116.2, m  $182^{\circ}$ ,  $185^{\circ}(\text{dec})$ . Possible impurity is *N,N'*-di-*tert*-butyl urea which is quite insoluble in  $H_2O$ . Recrystallise it from hot  $H_2O$ , filter off insoluble material, and cool from  $0^{\circ}$  to  $-5^{\circ}$  with stirring. Dry in vacuum at room temperature over  $KOH$  or  $H_2SO_4$ . If dried at higher temperatures, it sublimes slowly. It can be recrystallised from EtOH as long white needles or from 95% aqueous EtOH as plates. During melting point determination the bath temperature has to be raised rapidly as the urea sublimes slowly above  $100^{\circ}$  at 760mm. [Smith & Emerson *Org Synth Coll Vol III* 151 1955, Beilstein 4 IV 665.]

**n-Butyl vinyl ether** [111-34-2] M 100.2, b  $93.3^{\circ}$ ,  $d_4^{20}$  0.775. After five washings with equal volumes of water to remove alcohols (made slightly alkaline with  $KOH$ ), the ether is dried with sodium and distilled under vacuum, taking the middle fraction [Coombes & Eley *J Chem Soc* 3700 1957]. Store it over  $KOH$ . [Beilstein 1 IV 2052.]

**2-Butyne** [503-17-3] M 54.1, b 0°/253mm,  $d_4^{20}$  0.693. Keep it over Na wire for 24 hours, then fractionally distil it under reduced pressure into a cooled receiver. [Beilstein 1 IV 971.]

**2-Butyne-1,4-diol** [110-65-6] M 86.1, m 54-57°, 56-58°, b 238°. Crystallise the diol from EtOAc. [Beilstein 1 IV 2687.]

**Butyramide** [514-35-5] M 87.1, m 115°, b 230°. Crystallise it from acetone, \*benzene, CCl<sub>4</sub>/petroleum ether, 20% EtOH or water. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub>, CaCl<sub>2</sub> or 99% H<sub>2</sub>SO<sub>4</sub>. [Beilstein 2 H 275, 2 I 122, 2 II, 251, 2 III 616, 2 IV 804.]

**n-Butyric acid** [107-92-6] M 88.1, f -5.3°, b 163.3°,  $d_4^{20}$  0.961,  $n_D^{25}$  1.396, pK<sup>25</sup> 2.82. Distil the acid, then mix it with KMnO<sub>4</sub> (20g/L), and fractionally redistil, discarding the first third of the distillate [Vogel *J Chem Soc* 1814 1948]. [Beilstein 2 IV 779.]

**n-Butyric anhydride** [106-31-0] M 158.2, b 198°,  $d_4^{20}$  0.968. Dry the anhydride by shaking it with P<sub>2</sub>O<sub>5</sub>, then distilling it. [Beilstein 2 IV 802.]

**γ-Butyrolactone** [96-48-0] M 86.1, b 83.8°/12mm,  $d_4^{20}$  1.124. Dry the lactone over anhydrous CaSO<sub>4</sub>, then fractionally distil it. Handle it in a fume cupboard due to its TOXICITY. [Beilstein 17 V 7.]

**Butyronitrile** [109-74-0] M 69.1, b 117.9°,  $d_4^{20}$  0.793,  $n_D^{20}$  1.3846,  $n_D^{30}$  1.37954. Treat it with conc HCl until the smell of the isonitrile had gone, then dry with K<sub>2</sub>CO<sub>3</sub> and fractionally distil [Turner *J Chem Soc* 1681 1956]. Alternatively, it is twice heated at 75° and stirred for several hours with a mixture of Na<sub>2</sub>CO<sub>3</sub> (7.7g) and KMnO<sub>4</sub> (11.5g) per L of butyronitrile. The mixture is cooled, then distilled. The middle fraction is dried over activated alumina. [Schoeller & Wiemann *J Am Chem Soc* 108 22 1986, Beilstein 2 IV 806.]

**Butyryl chloride (butanoyl chloride)** [141-75-3] M 106.6, f -89°, b 101-102°/atm,  $d_4^{20}$  1.026,  $n_D^{20}$  1.412. Check IR to see if there is a significant peak at 3000-3500 cm<sup>-1</sup> (br) for OH. If OH is present then reflux it with less than one mole equivalent of SOCl<sub>2</sub> for 1 hour and distil directly. The fraction boiling between 85-100° is then refractionated at atmospheric pressure. Keep all apparatus free from moisture and store the product in sealed glass ampoules under N<sub>2</sub>. LACHRYMATORY; handle in a good fume hood. [Hefferich & Schaeffer *Org Synth Coll Vol I* 147 1941, Beilstein 2 IV 803.]

**Capric acid (decanoic acid)** [334-48-5] M 172.3, m 31.5°, b 148°/11mm,  $d_4^{20}$  0.886,  $n_D^{25}$  1.424, pK<sub>Est</sub> ~4.9. The acid is best purified by conversion into its methyl ester, b 114.0°/15mm (using excess MeOH, in the presence of H<sub>2</sub>SO<sub>4</sub>). The H<sub>2</sub>SO<sub>4</sub> and MeOH are removed, the ester is distilled *in vacuo* through a 3ft column packed with glass helices. The acid is then obtained from the ester by saponification and vacuum distillation. [Trachtman & Miller *J Am Chem Soc* 84 4828 1962, Beilstein 2 IV 1041.]

**n-Caproamide (n-hexanamide)** [628-02-4] M 115.2, m 100°, 100.5°. Recrystallise the amide from hot water. [Beilstein 2 H 324, 2 I 141, 2 II 286, 2 III 732.]

**Caproic acid (hexanoic acid)** [142-62-1] M 116.2, b 205.4°,  $d_4^{20}$  0.925,  $n_D^{20}$  1.417, pK<sup>25</sup> 4.85. Dry the acid with MgSO<sub>4</sub> and fractionally distil it from CaSO<sub>4</sub>. [Beilstein 2 IV 917.]

**Capronitrile (hexanenitrile)** [628-73-9] M 97.2, m -80°, b 163.7°,  $n_D^{20}$  1.4069,  $n_D^{25}$  1.4048. Wash the nitrile twice with half-volumes of conc HCl, then with saturated aqueous NaHCO<sub>3</sub>, dry over MgSO<sub>4</sub>, filter and distil it. [Beilstein 2 H 324, 2 I 141, 2 II 286, 2 III 733, 2 IV 930.]

**Caprylonitrile (heptylcyanide)** [124-12-9] M 125.2, m -45°, b 198-200°/~760mm,  $d_4^{20}$  0.812,  $n_D^{20}$  1.420. Wash the nitrile twice with half-volumes of conc HCl, then with saturated aqueous NaHCO<sub>3</sub>, dry over MgSO<sub>4</sub>, filter and distil it. [Beilstein 2 H 349, 2 I 148, 2 II 303, 2 III 798, 2 IV 993.]

**Carbon Black** [1333-86-4] **M 12.0, d 1.887/ml, bulk d 0.056g/ml, surface area >200m<sup>2</sup>/g, av pore diameter 64Å.** Leach the carbon for 24 hours with 1:1 HCl to remove oil contamination, then wash it repeatedly with distilled water. Dry it in air, and elute for one day each with \*benzene and acetone. Again dry it in air at room temperature, then heat it in a vacuum for 24 hours at 600° to remove adsorbed gases. [Tamamushi & Tamaki *Trans Faraday Soc* **55** 1007 1959.]

**Carbon disulfide** See entry in “Inorganic Compounds”, Chapter 5.

**Carbon tetrabromide** [558-13-4] **M 331.7, m 92.5°.** Reactive bromide is removed from CBr<sub>4</sub> by refluxing with dilute aqueous Na<sub>2</sub>CO<sub>3</sub>, then steam distilling, crystallising from EtOH, and drying in the dark under vacuum. [Sharpe & Walker *J Chem Soc* 157 1962.] It can be sublimed at 70° and low pressure. **It must not be dried with sodium.** [Beilstein **1** IV 85.]

**Carbon tetrachloride** [56-23-5] **M 153.8, b 76.8°, d<sup>25</sup> 1.5842.** For many purposes, careful fractional distillation gives adequate purification. Carbon disulfide, if present, can be removed by shaking vigorously for several hours with saturated KOH, separating, and washing with water: this treatment is repeated. The CCl<sub>4</sub> is shaken with conc H<sub>2</sub>SO<sub>4</sub> until there is no further coloration, then washed with water, dried with CaCl<sub>2</sub> or MgSO<sub>4</sub> and distilled (from P<sub>2</sub>O<sub>5</sub> if desired). **It must not be dried with sodium.** An initial refluxing with mercury for 2 hours removes sulfides. Other purification steps include passage of dry CCl<sub>4</sub> through activated alumina, and distillation from KMnO<sub>4</sub>. Carbonyl containing impurities can be removed by percolation through a Celite column impregnated with 2,4-dinitrophenylhydrazine (DNPH), H<sub>3</sub>PO<sub>4</sub> and water. (Prepared by dissolving 0.5g DNPH in 6ml of 85% H<sub>3</sub>PO<sub>4</sub> by grinding together, then mixing with 4ml of distilled water and 10g Celite.) [Schwartz & Parks *Anal Chem* **33** 1396 1961]. Photochlorination of CCl<sub>4</sub> has also been used: CCl<sub>4</sub> to which a small amount of chlorine has been added is illuminated in a glass bottle (e.g. for 24 hours with a 200W tungsten lamp near it), and, after washing out the excess chlorine with 0.02M Na<sub>2</sub>SO<sub>3</sub>, the CCl<sub>4</sub> is washed with distilled water and distilled from P<sub>2</sub>O<sub>5</sub>. It can be dried by passing through 4A molecular sieves and distilled. Another purification procedure is to wash CCl<sub>4</sub> with aqueous NaOH, then repeatedly with water and N<sub>2</sub> gas is bubbled through the liquid for several hours. After drying over CaCl<sub>2</sub> it is percolated through silica gel and distilled under dry N<sub>2</sub> before use [Klassen & Ross *J Phys Chem* **91** 3664 1987]. [Beilstein **1** IV 56.]

**Rapid purification:** Distil, discarding the first 10% of distillate or until the distillate is clear. The distilled CCl<sub>4</sub> is then stored over 5A molecular sieves.

**Carbon tetrafluoride** [75-73-0] **M 88.0, b -15°.** Purify CF<sub>4</sub> by repeated passage over activated charcoal at solid-CO<sub>2</sub> temperatures. Traces of air are removed by evacuating while alternately freezing and melting. *Alternatively*, liquefy CF<sub>4</sub> by cooling in liquid air and then fractionally distil it under vacuum. (The chief impurity originally present is probably CF<sub>3</sub>Cl). Use brass equipment. It is non-flammable, but is TOXIC. [Beilstein **1** H 59, **1** I 8, **1** II 11, **1** III 35, **1** IV 26.]

**Carbon tetraiodide** [507-25-5] **M 519.6, m 168°(dec).** Sublime ICl<sub>4</sub> *in vacuo*. [Beilstein **1** H 74, **1** IV 98.]

**Carbonyl sulfide** See “Inorganic Compounds” in Chapter 5.

**Cerulenin (helicocerin, 2R,3S-2,3-epoxy-4-oxo-7E,10E-dodecadienamide)** [17397-89-6] **M 223.3, m 93-94°, 93-95°, b 120°/10<sup>-8</sup>mm, [α]<sub>D</sub><sup>16</sup> +63° (c 2, MeOH).** It forms white needles from \*C<sub>6</sub>H<sub>6</sub>. It has also been purified by repeated chromatography through Florisil and silica gel. It is soluble in EtOH, MeOH, \*C<sub>6</sub>H<sub>6</sub>, slightly soluble in H<sub>2</sub>O and petroleum ether. The *dl*-form has **m 40-42°** (from \*C<sub>6</sub>H<sub>6</sub>/hexane), and the *2R,3S-tetrahydrocerulenin* has **m 86-87°, [α]<sub>D</sub><sup>20</sup> +44.4** (c 0.25, MeOH after 24 hours). [Ohruai & Emato *Tetrahedron Lett* 2095 1978, Sneda et al. *Tetrahedron Lett* 2039 1979, Broeckman & Thomas *J Am Chem Soc* **99** 2805 1977, Jakubowski et al. *J Org Chem* **47** 1221 1982, Beilstein **18/2** V 201.]

**Cetyl acetate** [629-70-9] **M 284.5, m 18.3°, b 158°/2mm, 204°/18mm, d<sub>4</sub><sup>20</sup> 0.861.** Distil the ester in a vacuum twice, then recrystallise it several times from diethyl ether/MeOH. [Beilstein **2** H 136, **2** II 146, **2** III 265, **2** IV 171.]

**Cetyl alcohol (1-hexadecanol)** [36653-82-4] **M 242.5, m 49.3°**. Crystallise the alcohol from aqueous EtOH or from cyclohexane. *Alternatively*, purify it by zone refining. The purity can be checked by gas chromatography. [Beilstein 1 H 429, 1 I 219, 1 II 466, 1 III 1815, 1 IV 1876.]

**Cetylamine** [629-54-9] **M 255.4, m 104-104.5°, 106-107°, b 235-236°/12mm**. Crystallise the amide from thiophene-free \*benzene and dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. It is slightly soluble in EtOH, Me<sub>2</sub>CO, CHCl<sub>3</sub> and toluene but insoluble in H<sub>2</sub>O. [Beilstein 2 H 374, 2 II 341, 2 III 975, 2 IV 1182.]

**Cetylamine (1-hexadecylamine)** [143-27-1] **M 241.5, m 48°, b 162-165°/5.2mm, pK<sup>25</sup> 10.60**. Crystallise the base from thiophene-free \*benzene and dry under vacuum over P<sub>2</sub>O<sub>5</sub>. Store away from CO<sub>2</sub> [Beilstein 4 IV 818.]

**Cetylammonium chloride** [1602-97-7] **M 278.0, m 178°**. Crystallise the salt from MeOH. [Beilstein 4 IV 818.]

**Cetyl bromide (1-bromohexadecane)** [112-82-3] **M 305.4, m 15°, 16-18°, b 144.8°/1mm, 171.5-172°/5mm, 190°/11mm, 193-196°/14mm, d<sup>25</sup> 0.999, n<sub>D</sub><sup>20</sup> 1.461**. It is prepared by boiling hexadecan-1-ol with 48% aqueous HBr and 25% H<sub>2</sub>SO<sub>4</sub> for several hours, distilling and purifying by shaking the bromide with H<sub>2</sub>SO<sub>4</sub>, washing with H<sub>2</sub>O, drying with K<sub>2</sub>CO<sub>3</sub> and fractionally distilling it *in vacuo*. [Bezzi & Lanza *Gazzetta* 80 180 1950, Vogel *J Chem Soc* 146 637 1943, Heston et al. *J Am Chem Soc* 72 2071 1950, Beilstein 1 H 17, 1 I 68, 1 II 138, 1 III 559, 1 IV 542.]

**Cetyl ether (dihexadecyl ether)** [4113-12-6] **M 466.9, m 54°**. Distil the ether in a vacuum then crystallise it several times from MeOH/\*benzene. [Beilstein 1 H 430, 1 II 467, 1 III 1820, 1 IV 1878.]

**Cetyltrimethylammonium bromide (cetrimonium bromide, CTAB)** [57-09-0] **M 364.5, m 227-235°(dec)**. Crystallise it from EtOH, EtOH/\*benzene or from wet acetone after extracting twice with petroleum ether. Shake it with anhydrous diethyl ether, filter and dissolve it in a little hot MeOH. After cooling in the refrigerator, the precipitate is filtered off at room temperature and re-dissolved in MeOH. Anhydrous ether is added and, after warming to obtain a clear solution, it is cooled and the crystalline material is collected and dried *in vacuo*. [Dearden & Wooley *J Phys Chem* 91 2404 1987, Hakemi et al. *J Am Chem Soc* 91 120 1987, Beilstein 4 IV 819.]

**Cetyltrimethylammonium chloride** [112-02-7] **M 320.0**. Crystallise the chloride from acetone/ether mixture, EtOH/ether, or from MeOH. [Moss et al. *J Am Chem Soc* 109 4363 1987, Beilstein 4 IV 819.]

**Charcoal** [7440-44-0] **M 12.0, m ~3550°**. Charcoal (50g) is added to 1L of 6M HCl and boiled for 45 minutes. The supernatant is discarded, and the charcoal is boiled with two more lots of HCl, then with distilled water until the supernatant no longer gives a test for chloride ion. The charcoal (now phosphate-free) is filtered onto a sintered-glass funnel and dried in air at 120° for 24 hours. [Lippin et al. *J Am Chem Soc* 76 2871 1954.] The purification can be carried out using a Soxhlet extractor (without cartridge), allowing longer extraction times. Treatment with conc H<sub>2</sub>SO<sub>4</sub> instead of HCl has been used to remove reducing substances.

**Chimyl alcohol (1-O-n-hexadecylglycerol)** [(±) 506-03-6, 10550-58-0 (*chimyl alcohol*)] **M 316.5, m 64°**. Recrystallise it from hexane. [Beilstein 1 III 2322.]

**Chloral (trichloroacetaldehyde)** [75-87-6, 302-17-0 (*hydrate*)] **M 147.4, b 26°/35mm, 98°/760mm, pK<sup>25</sup> 10.04**. Distil chloral, then dry it by distilling through a heated column of CaSO<sub>4</sub>. It readily forms a *hydrate* **m 57°**, an *alcoholate* **m 47.5°** [515-83-3] and an *ammonia adduct* **m 72-74°** [507-47-1]. It is a sedative and a hypnotic. [Beilstein 1 H 616, 1 I 328, 1 II 467, 1 III 2663, 1 IV 3142 for anhydr, 1 IV 3143 for hydrate.]

**Chloralacetone chloroform (2,2,2-trichloro-1-[2,2,2-trichloro-1,1-dimethylethoxy]ethanol)** [512-47-0] **M 324.9, m 65°**. Crystallise it from \*benzene. It sublimes on careful heating and hydrolyses in cold H<sub>2</sub>SO<sub>4</sub> to chloral acetone and chloroform. [Beilstein 1 H 622.]

**Chloroacetaldehyde dimethyl acetal** [97-97-2] **M 124.6, m -34.4°, b 64°/23mm, 71-72°/35mm, d<sub>4</sub><sup>20</sup> 1.0172, n<sub>D</sub><sup>20</sup> 1.4175.** Purify the acetal by fractional distillation, preferably in a vacuum. [Melhotra *J Indian Chem Soc* 36 4405 1959, *Beilstein* 1 IV 3134.]

**α-Chloroacetamide** [79-07-2] **M 93.5, m 121°, b 224-225°/743mm.** Recrystallise the amide from acetone and dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* 2 IV 490.]

**Chloroacetic acid** [79-11-8] **M 94.5, m 62.8°, b 189°, pK<sup>25</sup> 2.87.** Crystallise the acid from CHCl<sub>3</sub>, CCl<sub>4</sub>, \*benzene or water. Dry it over P<sub>2</sub>O<sub>5</sub> or conc H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. Further purification is by distillation from MgSO<sub>4</sub>, and by fractional crystallisation from the melt. Store it under vacuum or under dry N<sub>2</sub>. [Bernasconi et al. *J Am Chem Soc* 107 3621 1985, *Beilstein* 2 IV 474.]

**Chloroacetic anhydride** [541-88-8] **M 171.0, m 46°, b 122-123°/20mm, 203°/760mm, d<sub>4</sub><sup>20</sup> 1.5494.** Crystallise it from \*benzene. [Eglinton et al. *J Chem Soc* 1860 1954, *Beilstein* 2 IV 487.]

**Chloroacetone** [78-95-5] **M 92.5, b 61°/50mm, 119°/763mm, d<sub>4</sub><sup>20</sup> 1.15.** Dissolve it in water and shake it repeatedly with small amounts of diethyl ether which extracts, preferentially, 1,1-dichloroacetone present as an impurity. The chloroacetone is then extracted from the aqueous phase using a large volume of diethyl ether, and distill at slightly reduced pressure. It is dried with CaCl<sub>2</sub> and stored at Dry-ice temperature. *Alternatively*, it is kept over CaSO<sub>4</sub> for several hours, distilled and stored over CaSO<sub>4</sub>. It is steam volatile. The 2,4-dinitrophenylhydrazone forms yellow needles from EtOH with **m 120° or 124°.** [*Beilstein* 1 IV 3215.] **LACHRYMATOR with toxic vapour.**

**Chloroacetonitrile** [107-14-2] **M 75.5, b 125°.** Reflux it with P<sub>2</sub>O<sub>5</sub> for one day, then distil it through a helices-packed column. Also purified by gas chromatography. [*Beilstein* 2 IV 492.] **LACHRYMATOR AND HIGHLY TOXIC.**

**2-Chlorobutane** [78-86-4] **M 92.6, b 68.5°, d<sub>4</sub><sup>20</sup> 0.873, n<sub>D</sub><sup>25</sup> 1.3945.** Purify it in the same way as *n*-butyl chloride. [*Beilstein* 1 IV 248.]

***N*-Chlorocarbonyl isocyanate** [27738-96-1] **M 105.5, m -68°, b 63.6°/atm, d<sub>4</sub><sup>20</sup> 1.310.** Fractionally distil it at 760mm using a 40cm column. **TOXIC vapour, use a good fume hood.** Store it dry, its IR (film) has  $\nu_{\max}$  at 2260 (NCO), 1818 (CO) and 1420 (NCO sym) cm<sup>-1</sup>. [Jäckl & Sundmeyer *Chem Ber* 106 1752 1975.]

**2-Chloroethanol (ethylene chlorohydrin)** [107-07-3] **M 80.5, b 51.0°/31mm, 128.6°/760mm, d<sub>4</sub><sup>20</sup> 1.201, n<sub>D</sub><sup>15</sup> 1.444.** Dry it with, then distil it from, CaSO<sub>4</sub> in the presence of a little Na<sub>2</sub>CO<sub>3</sub> to remove traces of acid. [*Beilstein* 1 IV 1372.]

**2-Chloroethyl bromide (1-bromo-2-chloroethane)** [107-04-0] **M 143.4, b 106-108°, d<sub>4</sub><sup>20</sup> 1.723, n<sub>D</sub><sup>20</sup> 1.490.** Wash it with conc H<sub>2</sub>SO<sub>4</sub>, water, 10% Na<sub>2</sub>CO<sub>3</sub> solution, and again with water, then dry with CaCl<sub>2</sub> and fractionally distil before use. [*Beilstein* 1 H 89, 1 I 28, 1 II 61, 1 III 179, 1 IV 155.]

**2-Chloroethyl chloroformate** [627-11-2] **M 143.0, b 52-54°/12mm, 153°/760mm, d<sub>4</sub><sup>18</sup> 1.3760, n<sub>D</sub><sup>20</sup> 1.446.** Purify it by fractional distillation, preferably in a vacuum, and store it in a dry atmosphere. [Jones *J Chem Soc* 2735 1957, *Beilstein* 3 IV 24.]

**2-Chloroethyl vinyl ether** [110-75-8] **M 106.6, b 109°/760mm, d<sub>4</sub><sup>20</sup> 1.048, n<sub>D</sub><sup>20</sup> 1.437.** Wash the ether repeatedly with equal volumes of water made slightly alkaline with KOH, dry with sodium, and distil it under a vacuum. Stabilise it with ~0.01% of triethanolamine. [*Beilstein* 1 IV 2051.] **TOXIC.**

**Chlorofluoroacetonitrile (ClFCHCN)** [ $\pm$  92484-61-2 and 359-05-7] **M 93.4, b 66°/atm, d<sub>4</sub><sup>20</sup> 1.267, n<sub>D</sub><sup>25</sup> 1.3627.** It is prepared from ethyl chlorofluoroacetate (see below) by conversion to the amide (0.5mol, b 72-77°/1mm) which is added slowly to, P<sub>2</sub>O<sub>5</sub> (0.25mol), and the nitrile (3.5g) is distilled off at 50-70°/atmospheric and redistilled (30g, 66°/atmospheric). It is slowly hydrolysed by H<sub>2</sub>O but rapidly by aqueous alkali. [Young &

Tarrant *J Am Chem Soc* **71** 2432 1949, *Beilstein* **2** III 454]. Theoretical studies of the molecular dynamics of the enantiomers {(+)- [85196-70-9] and (-)- [85196-71-0]} have been made although they had not been separated [Roselli et al. *J Mol Liq* **28** 1 1984, Evans *Phys Rev A* **30** 2062 1984, Craig & Elsam *Chem Phys* **73** 349 1982]. Samarium (II) iodide [32248-43-4] promotes the reaction of chlorofluoroacetonitrile with aldehydes to form cyanofluorohydrins in the presence of HMPA [Asano et al. *Synthesis* 1309 2007].

**Chloroform** [67-66-3] **M 119.4, b 61.2°, d<sup>15</sup> 1.49845, d<sup>10</sup> 1.47060, n<sup>15</sup> 1.44858**. It reacts slowly with oxygen, or oxidising agents, when exposed to air and light, giving, mainly, phosgene, Cl<sub>2</sub> and HCl. Commercial CHCl<sub>3</sub> is usually stabilised with up to 1% of EtOH or of dimethylaminoazobenzene. Simplest purifications involve washing with water to remove the EtOH, drying with K<sub>2</sub>CO<sub>3</sub> or CaCl<sub>2</sub>, refluxing with P<sub>2</sub>O<sub>5</sub>, CaCl<sub>2</sub>, CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, and distilling. **It must not be dried with sodium**. The distilled CHCl<sub>3</sub> should be stored in the dark to avoid photochemical formation of phosgene. In an alternative purification, CHCl<sub>3</sub> (500ml) was shaken (mechanically) with several small portions of 12% H<sub>2</sub>SO<sub>4</sub> for 1 hour, washed thoroughly with water, saturated NaHCO<sub>3</sub>, washed again with water, and dried over CaCl<sub>2</sub> or K<sub>2</sub>CO<sub>3</sub> (100g) for 1 hour before filtering and distilling. After further drying for a short time over P<sub>2</sub>O<sub>5</sub>, the CHCl<sub>3</sub> was redistilled and stored over Drierite in the dark [Reynolds & Evans *J Am Chem Soc* **60** 2559 1938].

EtOH can be removed from CHCl<sub>3</sub> by passage through a column of activated alumina, or through a column of silica gel 4-ft long by 1.75-in diameter at a flow rate of 3ml/minute. (The alumina column, which can hold about 8% of its weight of EtOH, is regenerated by air drying and then heating at 600° for 6 hours. It is pre-purified by washing with CHCl<sub>3</sub>, then EtOH, leaving in conc H<sub>2</sub>SO<sub>4</sub> for about 8 hours, washing with water until the washings are neutral, then air drying, followed by activation at 600° for 6 hours. Just before use it is reheated for 2 hours at 154°.) [McLaughlin et al. *Anal Chem* **30** 1517 1958.]

Carbonyl-containing impurities can be removed from CHCl<sub>3</sub> by percolation through a Celite column impregnated with 2,4-dinitrophenylhydrazine (DNPH), phosphoric acid and water. (Prepared by dissolving 0.5g DNPH in 6ml of 85% H<sub>3</sub>PO<sub>4</sub> by grinding together, then mixing with 4ml of distilled water and 10g of Celite.) [Schwartz & Parks *Anal Chem* **33** 1396 1961]. Chloroform can be dried by distillation from powdered type 4A Linde molecular sieves. For use as a solvent in IR spectroscopy, chloroform is washed with water (to remove EtOH), then dried for several hours over anhydrous CaCl<sub>2</sub> and fractionally distilled. This treatment removes material absorbing near 1600 cm<sup>-1</sup>. (Percolation through activated alumina increases this absorbing impurity). [Goodspeed & Millson *Chem Ind (London)* 1594 1967, *Beilstein* **1** IV 42.]

**Rapid purification:** Pass through a column of basic alumina (Grade I, 10g/ml of CHCl<sub>3</sub>), and either dry by standing over 4A molecular sieves, or *alternatively*, distil from P<sub>2</sub>O<sub>5</sub> (3% w/v). Store away from light (to avoid formation of phosgene which is tested by shaking with conc NH<sub>3</sub> forming urea) and use as soon as possible.

**Chloromethyl methyl ether (MOMCl)** [107-30-2] **M 80.5, b 55-57°, d<sub>4</sub><sup>20</sup> 1.060, n<sub>D</sub><sup>20</sup> 1.396**. If suspect (check IR), shake it with saturated aqueous CaCl<sub>2</sub> solution, dry over CaCl<sub>2</sub> and fractionally distil taking the middle fraction. *Alternatively*, it can be prepared afresh. While working in a fume cupboard, a rapid stream of HCl gas ([7647-73-6] from a cylinder, or generated from concentrated H<sub>2</sub>SO<sub>4</sub> on NaCl) is bubbled through a solution of MeOH (438ml, 10.9moles) and technical formalin (900ml, 252g, 8.4moles, [50-00-0]) which is cooled with running water. After 2 hours two layers are formed, but HCl flow is continued for 2 to 3 hours further until saturation. The MOMCl layer is separated, the aqueous phase is saturated with CaCl<sub>2</sub>, and the additional MOMCl that is salted out is combined with the first lot, dried over CaCl<sub>2</sub>, filtered and fractionally distilled at atmospheric pressure. The fraction that distils at 55-60° is redistilled to give pure MOMCl (586-600g, 86-89%). [Marvel & Porter *Org Synth Coll Vol I* 377 1941, *Beilstein* **1** H 580, **1** I 304, **1** II 645, **1** III 2587, **1** IV 3046.] **VERY TOXIC VAPOUR and CARCINOGENIC**.

**1-Chloro-2-nitroethane** [625-47-8] **M 109.5, b 37-38°/20mm, 55°/60mm, 127.5°/atm, n<sub>D</sub><sup>20</sup> 1.4224, n<sub>D</sub><sup>25</sup> 1.4235**. Dissolve it in alkali, extract with ether (discard), then the aqueous phase is acidified with hydroxylamine hydrochloride, and the nitro compound is collected and fractionally distilled under reduced pressure. [Pearson & Dillon *J Am Chem Soc* **75** 2439 1953, *Beilstein* **1** H 101, **1** III 202, **1** IV 173.]

**Chloropicrin (trichloronitromethane)** [76-06-2] **M 164.5, b 112°**. Dry it over MgSO<sub>4</sub> and fractionally distil. [*Beilstein* **1** IV 106.] **EXTREMELY NEUROTOXIC, use appropriate precautions**.

**RS-2-Chloropropionic acid** [598-78-7] **M 108.5, b 98°/3mm, d<sub>4</sub><sup>20</sup> 1.182, n<sub>D</sub><sup>20</sup> 1.453 pK<sup>25</sup> 2.89.** Dry it with P<sub>2</sub>O<sub>5</sub> and fractionally distil it under vacuum. [Beilstein 2 IV 745.]

**S-(–)-2-Chloropropionic acid** [29617-66-1] **M 108.5, b 77°/10mm, 80.7°/10mm, 185-188°/atm, d<sub>4</sub><sup>25</sup> 1.2485, n<sub>D</sub><sup>25</sup> 1.436, [α]<sub>D</sub><sup>25</sup> –14.6° (neat).** Purify the acid by fractionating twice through a 115cm Podbielniak column (calculated 50 theoretical plates at atmospheric pressure) using a take-off ratio of 1:5. The *acid chloride* is prepared by dissolving the acid in SOCl<sub>2</sub>, adding a few drops of PCl<sub>3</sub>, refluxing and then distilling through a 30cm column, **b 53°/100mm, [α]<sub>D</sub><sup>25</sup> –4.6° (neat), d<sub>4</sub><sup>25</sup> 1.2689, n<sub>D</sub><sup>25</sup> 1.4368.** [Fu et al. *J Am Chem Soc* 76 6954 1954, Beilstein 2 IV 745.]

**3-Chloropropionic acid** [107-94-8] **M 108.5, m 41°, pK<sup>25</sup> 4.08.** Crystallise the acid from petroleum ether or \*benzene. [Beilstein 2 IV 748.]

**3-Chloropropyl bromide (1-bromo-3-chloropropane)** [109-70-6] **M 157.5, b 142-145°, n<sup>25</sup> 1.4732.** Wash it with conc H<sub>2</sub>SO<sub>4</sub>, water, 10% Na<sub>2</sub>CO<sub>3</sub> solution, water again and then dry with CaCl<sub>2</sub> and fractionally distil it just before use [Akagi et al. *J Am Chem Soc* 78 4034 1956]. [Beilstein 1 H 109, 1 IV 212.]

**2-Chloro-1,1,1-triethoxyethane** [51076-95-0] **M 196.7, b 75-80°/5mm, 91°/25mm, 147°/atm, d<sub>4</sub><sup>25</sup> 1.024, n<sub>D</sub><sup>20</sup> 1.422.** Distil the chloroethane in a vacuum; but if it is discoloured or suspect, then it is better to prepare it anew. To (EtO)<sub>3</sub>CH (ethyl orthoformate see [122-51-0]) in CCl<sub>4</sub> add a slight excess of *N*-chlorosuccinimide and heat to 60° while exposing to a sun lamp. An exothermic reaction occurs during 30 minutes; filter off the succinimide formed, evaporate and distil the residue *in vacuo*. [McElvain & Nelson *J Am Chem Soc* 64 1825 1942, Mylari et al. *Synth Commun* 19 2921 1989, Ueno et al. *J Med Chem* 34 2468 1991, Beilstein 2 III 442, 2 IV 482.]

**2-Chlorotriethylamine hydrochloride** [869-24-9] **M 172.1, m 208-210°, pK<sub>Est</sub> ~8.6 (free base).** Crystallise the salt from absolute MeOH (to remove highly coloured impurities). [Beilstein 4 III 240.]

**Chlorotrifluoroethylene (CTFE)** [79-38-9] **M 116.5, b –26 to –24°.** Scrub it with 10% KOH solution, then 10% H<sub>2</sub>SO<sub>4</sub> solution to remove inhibitors, dry and pass it through silica gel. It is stabilised with ~1% tributylamine. Use brass equipment. [Beilstein 1 III 646.] **TOXIC GAS.**

**Chlorotrifluoromethane** [75-72-9] **M 104.5, m –180°, b –81.5°.** Main impurities are CO<sub>2</sub>, O<sub>2</sub>, and N<sub>2</sub>. The CO<sub>2</sub> is removed by passage through saturated aqueous KOH, followed by concentrated H<sub>2</sub>SO<sub>4</sub>. The O<sub>2</sub> is removed using a tower packed with activated copper on Kieselguhr at 200°, and the gas is dried over P<sub>2</sub>O<sub>5</sub>. [Miller & Smyth *J Am Chem Soc* 79 20 1957, Beilstein 1 III 42, 1 IV 34.] **TOXIC GAS.**

**Choline acetate ([2-hydroxyethyl]trimethylammonium acetate)** [16586-35-7] **M 163.2, m varies with amount of H<sub>2</sub>O.** Choline acetate is a *very hygroscopic* solid and should be kept in well stoppered containers, preferably under N<sub>2</sub> or Ar. [Sasaki & Kobayashi *J Agric Chem Soc* 23 456 1949, *Chem Abstr* 47 2696 1953.] It forms a *gold complex* [HOCH<sub>2</sub>CH<sub>2</sub>N<sup>+</sup>(Me)<sub>3</sub>][AcO<sup>–</sup>][AuCl<sub>3</sub>] with **m 263°.** *Choline Reineckate* has **m 267-270°** (from aqueous Me<sub>2</sub>CO, **m 226-258°** was also reported). *Choline picrate* has **m 247-247.5°** (from EtOH, **m**'s of 237-238° and 242-245° were also reported) [Taylor & Kraus *J Am Chem Soc* 69 1732 1947] and *choline picrolonate* has **m 186°** (from EtOH). [Beilstein 4 IV 1444.]

**Choline chloride ([2-hydroxyethyl]trimethylammonium chloride)** [67-48-1] **M 139.6, m 302-305°(dec).** *Extremely deliquescent.* Check purity by AgNO<sub>3</sub> titration or by titration of OH-base after passage through an anion-exchange column. Crystallise it from absolute EtOH, or EtOH/Et<sub>2</sub>O, dry it under vacuum and store it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or Mg(ClO<sub>4</sub>)<sub>2</sub>. [Beilstein 4 IV 1443.]

**Citraconic acid (methylmaleic acid)** [498-23-7] **M 130.1, m 91°, pK<sub>1</sub><sup>25</sup> 2.2, pK<sub>2</sub><sup>25</sup> 5.60 (cis).** Steam distil and crystallise it from EtOH/ligroin. [Beilstein 2 H 768, 2 I 309, 2 II 652, 2 III 1938, 2 IV 2230.]

**Citraconic anhydride (methylmaleic anhydride)** [616-02-4] **M 112.1, m 8-9°, b 109°/30mm, 132°/74mm, 213°/760mm, d<sub>4</sub><sup>20</sup> 1.245, n<sub>D</sub><sup>20</sup> 1.472.** Possible contamination is from the acid formed by hydrolysis. If the IR

has OH bands, then reflux with  $\text{Ac}_2\text{O}$  for 30 minutes, evaporate, and distil the residue in a vacuum; otherwise distil in a vacuum. Store it in a dry atmosphere. [Vaughan & Andersen *J Am Chem Soc* **77** 6702 1955, Vaughan & Andersen *J Org Chem* **21** 680 1956, *Beilstein* **17** H 440, **17** I 234, **17** II 448, **17** III/IV 5912, **17/11** V 65.]

**Citric acid ( $\text{H}_2\text{O}$ )** [5949-29-1 (monohydrate); 77-92-9 (anhydrous)] **M 210.1, m 156-157°, 153° (anhyd), pK<sub>1</sub><sup>25</sup> 2.96, pK<sub>2</sub><sup>25</sup> 4.38, pK<sub>3</sub><sup>25</sup> 5.68.** Crystallise it from hot  $\text{H}_2\text{O}$  solution (w/w solubility is 54% at 10°, 71% at 50° and 84% at 100°). The *monohydrate* (softens at ~75° and melts at ~100°) dehydrates in air or when heated gently above 40°. The *triethylester* ([77-93-0]) **M 276.3, b 127°/1mm, 294°/atm, d<sub>4</sub><sup>20</sup> 1.137, n<sub>D</sub><sup>20</sup> 1.4420** is a bitter tasting oil. [*Beilstein* **3** H 556 and 568, **3** IV 1272.]

**Citronellal (3,7-dimethyloctan-6-al)** [*R*(+) 2385-77-5, *S*(-) 5949-05-3] **M 154.3, b 67°/4mm, 89°/14mm, 104-105°/21mm, 207°/760mm, [α]<sub>546</sub><sup>20</sup> (+) and (-) 20°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 16.5° (neat).** Fractionally distil it. Alternatively, extract it with  $\text{NaHSO}_3$  solution, wash it with  $\text{Et}_2\text{O}$ , then acidify it to decompose the bisulfite adduct and extract with  $\text{Et}_2\text{O}$ , dry ( $\text{Na}_2\text{SO}_4$ ), evaporate and distil. Check for purity by hydroxylamine titration. The ORD in MeOH (c 0.167) has: [α]<sub>700</sub><sup>+9°</sup>, [α]<sub>589</sub><sup>+11°</sup>, [α]<sub>275</sub><sup>+12°</sup> and [α]<sub>260</sub><sup>+12°</sup>. The *semicarbazone* has **m 85°**, and the *2,4-dinitrophenylhydrazone* has **m 79-80°**. [(+)-compound: Tietze & Beifuss *Org Synth* **71** 167 1993, IR: Carroll et al. *J Chem Soc* 3457 1950, ORD: Djerassi & Krakower *J Am Chem Soc* **81** 237 1959, *Beilstein* **1** IV 3515.]

**β-Citronellene (2,6-dimethylocta-2,7-diene)** [*S*(+) 2436-90-0, *R*(-) 10281-56-8] **M 138.3, b 153-154°/730mm, 155°/atm, d<sub>4</sub><sup>22</sup> 0.757, n<sub>D</sub><sup>22</sup> 1.431, [α]<sub>546</sub><sup>20</sup> (+) and (-) 13°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 10° (neat).** Purify it by distillation over Na (three times) and fractionation. [(-) Arigoni & Jaeger *Helv Chim Acta* **37** 881 1954, (+) Eschenmoser & Schinz *Helv Chim Acta* **33** 171 1950, *Beilstein* **1** IV 1059-1060.]

**β-Citronellol (3,7-dimethyloctan-6-ol)** [*R*(+): 11171-61-9, *S*(-): 106-22-9] **M 156.3, b 47°/1mm, 102-104(110)°/10mm, 112-113°/12mm, 221-224°/atm, 225-226°/atm, d<sub>4</sub><sup>24</sup> 0.8551, n<sub>D</sub><sup>24</sup> 1.4562, [α]<sub>546</sub><sup>20</sup> (+) and (-) 6.3°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 5.4° (neat).** Purify them by distillation through a cannon packed (Ni) column and the main cut collected at 84°/14mm and redistilled. Also purify *via* the benzoate. [IR: Eschenazi *J Org Chem* **26** 3072 1961, Naves *Bull Soc Chim Fr* 505 1951, *Beilstein* **1** IV 2188.]

**Crotonaldehyde (2-butenal)** [123-73-9] **M 70.1, b 104-105°, d<sub>4</sub><sup>20</sup> 0.851, n<sub>D</sub><sup>20</sup> 1.437.** Fractionally distil it under  $\text{N}_2$ , through a short Vigreux column. Stabilise it with 0.01% of 2,6-di-tert-butyl-*p*-cresol and store it in sealed ampoules. [*Beilstein* **1** IV 3447.]

**trans-Crotonic acid (trans-2-butenic acid)** [107-93-7] **M 86.1, m 72-72.5°, pK<sub>1</sub><sup>25</sup> -6.17 (aqueous  $\text{H}_2\text{SO}_4$ ), pK<sub>2</sub><sup>18</sup> 4.71.** Distil the acid under reduced pressure and/or recrystallise it from petroleum ether (b 60-80°) or water, or by partial freezing of the melt. [*Beilstein* **2** IV 1498.]

***E*- and *Z*-Crotonitrile (mixture)** [4786-20-3] **M 67.1, b 120-121°, d<sub>4</sub><sup>20</sup> 1.091, n<sub>D</sub><sup>20</sup> 1.4595.** Separate the mixture by preparative GLC on a column using 5% FFAP on Chromosorb G. [Lewis et al. *J Am Chem Soc* **108** 2818 1986, *Beilstein* **2** IV 1507.]

**trans-Crotonoyl chloride** [625-35-4] **M 104.5, b 42-45°/20mm, 124-125°/760mm, d<sub>4</sub><sup>20</sup> 1.080, n<sub>D</sub><sup>20</sup> 1.4570.** If the IR of a film has no OH bands, then fractionally distil it, taking the middle fraction and redistilling it. If OH bands are present then add excess of oxalyl chloride, reflux for 3 hours, then distil off the reagent and fractionally distil the crotonoyl chloride as before. Stabilise the distillate with 160ppm of hydroquinone. The *amide* forms needles **m 158°** from aqueous ammonia, and the *anilide* also forms needles from  $\text{H}_2\text{O}$ , but with **m 115-118°**. [*Beilstein* **2** H 411, **2** I 188, **2** II 392, **2** III 1265, **2** IV 1506.]

**Crotyl bromide** [29576-14-5] **M 135.0, b 103-105°/740mm, n<sub>D</sub><sup>25</sup> 1.4792.** Dry the bromide with  $\text{MgSO}_4/\text{CaCO}_3$  mixture and fractionally distil it through an all-glass Todd column. [*Beilstein* **1** IV 789.]

**Cyanoacetamide** [107-91-5] **M 84.1, m 119.4°.** Crystallise the amide from MeOH/dioxane (6:4), then water

and dry it over P<sub>2</sub>O<sub>5</sub> under vacuum. [*Beilstein* 2 IV 1891.]

**Cyanoacetic acid** [372-09-8] **M 85.1, m 70.9-71.1°**, **pK<sup>25</sup> 2.47**. Recrystallise the acid to constant melting point from \*benzene/acetone (2:3), and dry it over silica gel. [*Beilstein* 2 H 583, 2 I 253, 2 II 530, 2 III 1626, 2 IV 1888.]

**Cyanoacetic acid hydrazide** [140-87-4] **M 99.1, m 114.5-115°**. Crystallise the hydrazide from EtOH. The *hydrochloride* has **m 178-180°** and the *benzylidene* derivative has **m 178°**. It is converted to 3-*oxo-5-iminopyrazolidine* in hot 40% aqueous NaOH. [*Beilstein* 2 H 591, 2 I 256, 2 III 1636.] **IRRITANT**.

**Cyanoguanidine (dicyanodiamide)** [461-58-5] **M 84.1, m 209.5°**, **pK<sup>25</sup> -0.4**. Recrystallise cyanoguanidine from water or EtOH. [*Beilstein* 3 IV 160.]

***n*-Decane** [124-18-5] **M 142.3, m -29.7°**, **b 57.6°/10mm, 174.1°**, **d<sub>4</sub><sup>20</sup> 0.7300, n<sub>D</sub><sup>20</sup> 1.4102, n<sup>25</sup> 1.40967**. It can be purified by shaking with conc H<sub>2</sub>SO<sub>4</sub>, washing with water, aqueous NaHCO<sub>3</sub>, and more water, then drying with MgSO<sub>4</sub>, refluxing with Na and distilling. Also purify through a column of silica gel or alumina. It has been purified by azeotropic distillation with 2-butoxyethanol, the alcohol being washed out of the distillate using water, the decane is next dried and redistilled. It can be stored with NaH. Further purification can be achieved by preparative gas chromatography on a column packed with 30% SE-30 (General Electric methyl-silicone rubber) on 42/60 Chromosorb P at 150° and 40psig, using helium [Chu *J Chem Phys* 41 226 1964]. It is soluble in EtOH and Et<sub>2</sub>O. [*Beilstein* 1 IV 484.]

**Decan-1,10-diol** [112-47-0] **M 174.3, m 72.5-74°**. Crystallise the diol from dry ethylene dichloride. [*Beilstein* 1 IV 2613.]

***n*-Decanol (*n*-decyl alcohol)** [112-30-1] **M 158.3, f 6.0°**, **b 109°/8mm, 231°/atm, d<sub>4</sub><sup>20</sup> 0.823, n<sub>D</sub><sup>20</sup> 1.434**. Fractionally distil *n*-decanol in an all-glass unit at 10mm pressure (**b 110°**), then fractionally crystallise by partial freezing. Also purify by preparative GLC, and by passage through alumina before use. [*Beilstein* 1 IV 1815.]

***n*-Decyl bromide (1-bromodecane)** [112-29-8] **M 221.2, b 117-118°/15.5mm, d<sub>4</sub><sup>20</sup> 1.066**. Shake the it with H<sub>2</sub>SO<sub>4</sub>, wash with H<sub>2</sub>O, dry with K<sub>2</sub>CO<sub>3</sub>, and fractionally distil it. [*Beilstein* 1 IV 470.]

**Decyltrimethylammonium bromide** [2082-84-0] **M 280.3, m 239-242°**. Crystallise the salt from 50% (v/v) EtOH/Et<sub>2</sub>O, or from acetone and wash with ether. Dry it under vacuum at 60°. Also recrystallise it from EtOH and dry it over silica gel. [McDonnell & Kraus *J Am Chem Soc* 73 2170 1952, Dearden & Wooley *J Phys Chem* 91 2404 1987, *Beilstein* 4 IV 784.]

**Diacetamide** [625-77-4] **M 101.1, m 75.5-76.5°**, **b 222-223°**. Purify the amide by recrystallisation from MeOH [Arnett & Harrelson *J Am Chem Soc* 109 809 1987]. [*Beilstein* 2 H 181.]

**(+)-Di-*O*-acetyl-*L*-tartaric anhydride [(*R,R*)-2,3-diacetoxysuccinic anhydride]** [6283-74-5] **M 216.2, m 129-132°, 133-134°, 135°, 137.5°**, **[α]<sub>D</sub><sup>20</sup> +97.2° (c 0.5, dry CHCl<sub>3</sub>)**, **[α]<sub>D</sub><sup>20</sup> +60° (c 6, Me<sub>2</sub>CO)**. If the IR is good, i.e. no OH bands, then keep it in a vacuum desiccator overnight (over P<sub>2</sub>O<sub>5</sub>/paraffin) before use. If OH bands are present then reflux 4g in Ac<sub>2</sub>O (12.6ml) containing a few drops of conc H<sub>2</sub>SO<sub>4</sub> for 10 minutes (use a relatively large flask), pour onto ice, collect the crystals, wash with dry \*C<sub>6</sub>H<sub>6</sub> (2 x 2ml), stir with 17ml of cold Et<sub>2</sub>O, filter and dry in it a vacuum desiccator as above, or store it in dark evacuated ampoules under N<sub>2</sub> in small aliquots. It is not very stable in air, the melting point of the crystals drop one degree in the first four days then remains constant (132-134°). If placed in a stoppered bottle, it becomes gummy and the **m** falls 100° in three days. Recrystallisation leads to decomposition. If good quality anhydride is required it, should be prepared freshly from tartaric acid. It sublimes in a CO<sub>2</sub> atmosphere. [Shriner & Furrow *Org Synth Coll Vol* IV 242 1963, Bell *Aust J Chem* 34 671 1981, *Beilstein* 18 III/IV 2296.]

**Diallyl amine** (*N*-2-propenyl-2-propen-1-amine) [124-02-7] **M 97.2, b 107-111°/760mm, 112°/760mm, d<sub>4</sub><sup>20</sup> 0.789, n<sub>D</sub><sup>20</sup> 1.440, pK<sup>20</sup> 9.42.** Keep the amine over KOH pellets overnight, decant and distil it from a few pellets of KOH at atmospheric pressure (b 108-111°), then fractionate through a Vigreux column. [Vliet *J Am Chem Soc* **46** 1307 1924, *Org Synth Coll Vol* **1** 201 1941.] The *hydrochloride* has **m 164-165°** (from Me<sub>2</sub>CO/EtOH). [Butler & Angels *J Am Chem Soc* **79** 3128 1957.]

**(+)-*N,N'*-Diallyl tartardiamide (DATD)** [58477-85-3] **M 228.3, m 184°, [α]<sub>546</sub> +141° (c 3, MeOH).** Wash DATD with Et<sub>2</sub>O containing 10% EtOH until the washings are clear and colourless, and dry *in vacuo*. [*FEBS Lett* **7** 293 1970, *Beilstein* **4** H 218.]

**1,4-Diaminobutane dihydrochloride (putrescine 2HCl)** [333-93-7] **M 161.1, m >290°, pK<sub>1</sub><sup>25</sup> 9.63, pK<sub>2</sub><sup>25</sup> 10.80.** Crystallise the salt from EtOH/H<sub>2</sub>O. [*Beilstein* **4** IV 1284.]

**2,2'-Diaminodiethylamine (diethylenetriamine)** [111-40-0] **M 103.2, b 208°, d<sub>4</sub><sup>20</sup> 0.95, n<sub>D</sub><sup>20</sup> 1.483, pK<sub>1</sub><sup>25</sup> 4.34, pK<sub>2</sub><sup>25</sup> 9.13, pK<sub>3</sub><sup>25</sup> 9.94.** Dry the amine with Na and distil, preferably under reduced pressure, or in a stream of N<sub>2</sub>. [*Beilstein* **4** IV 1284.]

§ Polymer-bound diethylenetriamine is commercially available.

**3,3'-Diaminodipropylamine (Norspermidine, bis-[3-aminopropyl]amine)** [56-18-8] **M 131.2, b 152°/50mm, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>20</sup> 1.481, pK<sub>1</sub><sup>25</sup> 7.72, pK<sub>2</sub><sup>25</sup> 9.57, pK<sub>3</sub><sup>25</sup> 10.65.** Dry the amine with Na and distil it under vacuum. [*Beilstein* **4** IV 1278.]

**1,8-Diaminooctane** [373-44-4] **M 144.3, m 50-52°, 51-52°, 52-53°, b 121°/18mm, 120°/24mm, pK<sub>1</sub><sup>20</sup> 10.1, pK<sub>2</sub><sup>20</sup> 11.0.** Distil the diamine under vacuum in an inert atmosphere (N<sub>2</sub> or Ar), cool and store the distillate in an inert atmosphere in the dark. The *dihydrochloride* has **m 273-274°**. [Nae & Le *Helv Chim Acta* **15** 55 1955, *Beilstein* **4** III 612.]

**1,5-Diaminopentane** [462-94-2] **M 102.2, m 14-16°, b 78-80°/12mm, 101-103°/35mm, 178-180°/750mm, d<sub>4</sub><sup>20</sup> 0.869, n<sub>D</sub><sup>20</sup> 1.458, pK<sub>1</sub><sup>20</sup> 10.02, pK<sub>2</sub><sup>20</sup> 10.96.** Purify the base by distillation, after standing over KOH pellets (at room temperature, i.e. liquid form). Its *dihydrochloride* has **m 275°** (sublimes in a vacuum), and its *tetraphenyl boronate* has **m 164°**. [Schwarzenbach et al. *Helv Chim Acta* **35** 2333 1952, *Beilstein* **4** IV 1310.]

**1,3-Diaminopropane dihydrochloride** [10517-44-9] **M 147.1, m 243°, pK<sub>1</sub><sup>25</sup> 8.29, pK<sub>2</sub><sup>25</sup> 10.30.** Crystallise the salt from EtOH/H<sub>2</sub>O. [*Beilstein* **4** IV 1258 free base.]

**1,3-Diaminopropan-2-ol** [616-29-5] **M 90.1, m 38-40°, pK<sub>1</sub><sup>25</sup> 7.94, pK<sub>2</sub><sup>25</sup> 9.57.** Dissolve it in an equal amount of water, shake it with charcoal and distil it at 68°/0.1mm. The distillate solidifies. It is too viscous to be distilled through a packed column. [*Beilstein* **4** IV 1694.]

**1,11-Diamino-3,6,9-trioxaundecane [2,2'-oxybis(ethyleneoxy)bisethylamine]** [929-75-9] **M 192.3, b 115°/0-2mm, 138-153°/1mm, 145-150°/2mm, 135°/3mm, n<sub>D</sub><sup>20</sup> 1.440.** Distil the diamine under vacuum, but if it is suspect then dissolve it in EtOH add 6N HCl boil for 1 hour, evaporate to dryness, and dry the residue in a vacuum. The solid di-hydrochloride residue is treated with 4N NaOH to release the free base which is extracted into CH<sub>2</sub>Cl<sub>2</sub>, the extract is dried over K<sub>2</sub>CO<sub>3</sub>, filtered, evaporated and then distilled in a vacuum. It is a hydrophilic linker in dendrimer formation or biotinylation. [Zhan et al. *J Am Chem Soc* **123** 8914 2001, Han et al. *Synth Commun* **21** 79 1999, Bogatiskii et al. *J Org Chem USSR* (English Transl) **16** 1124 1980, McReynolds et al. *Bioorg Med Chem* **10** 625 2002.]

**Diazomethane [H<sub>2</sub>CN<sub>2</sub>]** [334-88-3] **M 42.0, m -145°, b -23°.** Diazomethane is produced when *N*-methyl-*N*-nitroso-amides, e.g. *N*-methyl-*N*-nitroso-urea, *N,N'*-dimethyl-*N,N'*-dinitroso-oxamide, *bis*-(*N*-methyl-*N*-nitroso)terephthalamide or *p*-toluenesulfonyl-*N*-methyl-*N*-nitrosoamide (Diazald), are treated with strong solutions of NaOH or KOH. [Fieser and Fieser's *Reagents for Organic Synthesis* **1** 191 1967, see also **2** 102 1969.] It has been commonly prepared from Diazald as it is commercially available, although *N*-methyl-*N*-nitroso-urea, which is readily prepared and stored below 5° [Arndt *Org Synth Coll Vol* **II** 461 1943], is also

frequently used. *Diazomethane is a sweet smelling, highly irritating yellow gas which is TOXIC when inhaled, causing severe irritation, pulmonary oedema, asthma, chest pains, headaches, weakness as well as producing hypersensitivity reactions and skin irritation. It is a carcinogen. Diazomethane and its precursors should be handled in an efficient fume cupboard; and because of its potentially explosive nature experiments should be carried out behind a safety screen with face and eye protection.*

**Preparation:** A distilling flask (100ml) containing KOH (5g) in H<sub>2</sub>O (8ml) and diluted with 95% EtOH (25ml) is heated to ~65° using a water bath, and a solution of Diazald (21.4g, see [80-11-5]) in Et<sub>2</sub>O (130ml) is added through a dropping funnel in ~25 minutes while the ether which is yellow in colour due to the diazomethane distils off, and the distillate is collected in a receiver immersed in ice. When addition of Diazald is complete, a further volume of Et<sub>2</sub>O (20ml) is added dropwise and distillation is continued until the Et<sub>2</sub>O distillate is colourless. The yield of diazomethane in the distillate is ~3g. [Note that if 95% EtOH is not added to the KOH solution no diazomethane is formed since the base does not dissolve into the Et<sub>2</sub>O.] Do **NOT** redistil this ethereal solution as an explosion may result, particularly from the smaller volume left in the flask. [For alcohol-free diazomethane distillate, the distilling flask should contain diethyleneglycol monomethyl ether (35ml), Et<sub>2</sub>O (10ml) and KOH (6g) in H<sub>2</sub>O (10ml) to which is added the ethereal Diazald as before.] [deBoer & Backer *Org Synth* **34** 96 1954, *Org Synth Coll Vol VI* 943 1963, *Rec Trav Chim Pays Bas* **73** 229 1954.] By using <sup>13</sup>C, <sup>14</sup>C and/or <sup>2</sup>H or <sup>3</sup>H methyl labeled Diazald (some of these are also commercially available) appropriately labeled diazomethane can be prepared. The amount of diazomethane is determined by adding a measured volume of 0.2N pure benzoic acid in ether to an aliquot of the ethereal diazomethane solution which will discharge the yellow colour, meaning that all the diazomethane has been consumed. Water is added to this mixture and the excess of benzoic acid is then titrated with 0.1N aqueous NaOH. (See the safer Me<sub>3</sub>SiCHN<sub>2</sub> [18107-18-1].) For most purposes it is not necessary to distil the diazomethane in ether. In such cases it is better to use *N*-methyl-*N*-nitrosourea because the unwanted products of the reaction are soluble in water. To a mixture of 50% aqueous KOH (60ml) and pure Et<sub>2</sub>O (200ml) at 0 to 5° is added solid nitrosomethylurea (20g) with shaking or stirring. The ethereal layer becomes deep yellow and is decanted or siphoned off; or separated if in a separating funnel. The aqueous layer can be extracted with Et<sub>2</sub>O to get a further amount of diazomethane. The ethereal solutions can be dried over KOH pellets and can be used directly or kept at ~0° for 2-3 days. [Arndt *Org Synth Coll Vol II* 165 1943] Diazomethane is a very powerful methylating agent and, in so doing the yellow colour disappears and N<sub>2</sub> bubbles off. It is a versatile reagent in organic synthesis [Black *Aldrichim Acta* **16** 2 1983.]

**Di-*O*-benzoyl-(*R* and *S*)-tartaric acid (H<sub>2</sub>O)** [*R*-(+) 17026-42-5, *S*-(-) 2743-38-6 (*anhydrous*), 62708-56-9 (*hydrate*)] **M 376.3, m 88-89° (hydrate), 173° (anhydrous), [α]<sub>D</sub><sup>20</sup> (+) and (-) 136° (c 2, EtOH), [α]<sub>D</sub><sup>20</sup> (+) and (-) 117° (c 5, EtOH), pK<sub>Est(1)</sub> ~2.9, pK<sub>Est(2)</sub> ~4.2.** Crystallise the acid from water (18g from 400 ml boiling H<sub>2</sub>O) and stir vigorously while cooling in order to obtain crystals; otherwise an oil will separate which solidifies on cooling. Dry it in a vacuum desiccator over KOH/H<sub>2</sub>SO<sub>4</sub> (yield 16.4g) as *monohydrate*, **m 88-89°**. It crystallises from xylene as the *anhydrous acid*, **m 173° (150-153°)**. It does not crystallise from \*C<sub>6</sub>H<sub>6</sub>, toluene, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (oil), or CHCl<sub>3</sub>/petroleum ether. [Butler & Cretcher *J Am Chem Soc* **55** 2605 1933, Acs et al. *Tetrahedron* **41** 2465 1985, *R*(+) *Beilstein* **9** IV 557, *S*(-) *Beilstein* **9** III 870.]

***trans*-1,4-Dibromobut-2-ene** [821-06-7] **M 213.9, m 54°, b 85°/10mm.** Crystallise the dibromide from ligroin and/or distil it in a vacuum. [*Beilstein* **1** IV 79.]

**Dibromodichloromethane** [594-18-3] **M 242.7, m 21.8°, b 66°/81mm, d<sub>4</sub><sup>20</sup> 2.433, n<sub>D</sub><sup>25</sup> 1.5499.** Crystallise CBr<sub>2</sub>Cl<sub>2</sub>, repeatedly from its melt, after washing with aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> and drying with BaO. *Alternatively*, distil it in a vacuum. Store away from light. [*Beilstein* **1** H 68, **1** III 88, **1** IV 82.]

**1,2-Dibromoethane** [106-93-4] **M 187.9, f 10.0°, b 29.1°/10mm, 131.7°/760mm, d 2.179, n<sub>D</sub><sup>15</sup> 1.54160.** Wash the dibromide with conc HCl or H<sub>2</sub>SO<sub>4</sub>, then water, aqueous NaHCO<sub>3</sub> or Na<sub>2</sub>CO<sub>3</sub>, more water, and dry it with CaCl<sub>2</sub>. Fractionally distil it. *Alternatively*, keep in daylight with excess bromine for 2 hours, then extract with aqueous Na<sub>2</sub>SO<sub>3</sub>, wash with water, dry with CaCl<sub>2</sub>, filter and distil. It can also be purified by fractional crystallisation by partial freezing. Store it in the dark. [*Beilstein* **1** H 90, **1** I 28, **1** II 61, **1** III 182, **1** IV 158.]

**Dibromomaleic acid** [608-37-7] **M 273.9, m 123.5°, 125°(dec), pK<sub>1</sub><sup>25</sup> 1.45, pK<sub>2</sub><sup>25</sup> 4.62.** It has been recrystallised from Et<sub>2</sub>O or Et<sub>2</sub>O/CHCl<sub>3</sub>. It is soluble in AcOH, also slightly soluble in water, but insoluble in

\*C<sub>6</sub>H<sub>6</sub> and CHCl<sub>3</sub>. [Salmony & Simonis *Chem Ber* **38** 2583 1905, Ruggli *Helv Chim Acta* **3** 566 1929, *Beilstein* **3** IV 2224.]

**1,3-Dibromopropane** [109-64-8] **M 201.9, f –34.4°, b 63-63.5°/26mm, 76-77°/40mm, 90°/80mm, 165°/atm, d<sub>4</sub><sup>20</sup> 1.977, n<sub>D</sub><sup>20</sup> 1.522.** Wash the dibromide with dilute aqueous Na<sub>2</sub>CO<sub>3</sub>, then water. Dry and fractionally distil it under reduced pressure. [*Beilstein* **1** IV 216.]

**meso-2,3-Dibromosuccinic acid** [608-36-6] **M 275.9, m 288-290°(sealed tube, dec), pK<sub>I</sub><sup>20</sup> 1.56, pK<sub>2</sub><sup>20</sup> 2.71.** Crystallise the acid from distilled water, keeping the temperature below 70°. [*Beilstein* **2** IV 1930.]

**1,2-Dibromotetrafluoroethane** [124-73-2] **M 259.8, b 47.3°/760mm.** Wash it with water, then with weak alkali. Dry with CaCl<sub>2</sub> or H<sub>2</sub>SO<sub>4</sub> and distil it. [Locke et al. *J Am Chem Soc* **56** 1726 1934.] Also purify it by gas chromatography on a silicone DC-200 column.

**Di-*n*-butylamine** [111-92-2] **M 129.3, b 159°, d<sub>4</sub><sup>20</sup> 0.761, n<sub>D</sub><sup>20</sup> 1.41766, pK<sup>25</sup> 11.25.** Dry this strong base with LiAlH<sub>4</sub>, CaH<sub>2</sub> or KOH pellets, filter and distil it from BaO or CaH<sub>2</sub>. [*Beilstein* **4** IV 550.]

**Di-*tert*-butylazodicarboxylate** [870-50-8] **M 230.3, m 89-92°, 90-92°.** The *tert*-butyl ester has the advantage over the ethyl ester (below) in being a solid and more acid labile. It crystallises from ligroin and is best purified by covering the dry solid (22g) with petroleum ether (b 30-60°, 35-40 ml) heating to boiling and adding ligroin (b 60-90°) until the solid dissolves. On cooling, large lemon yellow crystals of the ester separate (~20g), **m 90.7-92°.** Evaporation of the filtrate gives a further crop of crystals [Carpino & Crowley *Org Synth* **44** 18 1964]. This reagent is useful in the Mitsunobu reaction [Mitsunobu *Synthesis* **1** 1981, Gennari et al. *J Am Chem Soc* **108** 6394 1986, Evans et al. *J Am Chem Soc* **108** 6394 1986, Hughes *Org React* **42** 335 1992, Dodge et al. *Org Synth* **73** 110 1996, Hughes *Org Prep Proc Int* **28** 127 1996, Ferguson & Marcelle *J Am Chem Soc* **128** 4576 2006, see also **DEAD** and **DIAD** below].

**Dibutylcarbitol (di[ethyleneglycol]-dibutyl ether, bis[2-butoxyethyl]-ether** [112-73-2] **M 218.3, b 125-130°/0.1mm, d<sub>4</sub><sup>20</sup> 0.883, n<sub>D</sub><sup>20</sup> 1.424.** Dibutylcarbitol is freed from peroxides by slow passage through a column of activated alumina. The eluate is then shaken with Na<sub>2</sub>CO<sub>3</sub> (to remove any remaining acidic impurities), washed with water, and stored with CaCl<sub>2</sub> in a dark bottle [Tuck *J Chem Soc* 3202 1957]. [*Beilstein* **1** IV 2395.]

**Di-*tert*-butyl dicarbonate (di-*tert*-butyl pyrocarbonate, Boc anhydride)** [24424-99-5] **M 218.3, m 23° (21-22°), b 55-56°/0.15mm, 62-65°/0.4mm, d<sub>4</sub><sup>20</sup> 0.950, n<sub>D</sub><sup>20</sup> 1.409.** Melt the ester by heating at ~35°, and distil it in a vacuum. If the IR and NMR (ν<sub>max</sub> 1810m and 1765 cm<sup>-1</sup>, δ in CCl<sub>4</sub> 1.50 singlet) suggest that it is very impure, then wash it with an equal volume of H<sub>2</sub>O containing citric acid to make the aqueous layer slightly acidic, collect the organic layer, dry it over anhydrous MgSO<sub>4</sub> and distil it in a vacuum. Store it away from moisture at ~4°. If some moisture enters the container pressure may develop since it hydrolyses to *tert*-BuOH and CO<sub>2</sub>. It is also available commercially as a 1.0 M solution in THF. It is **FLAMMABLE**. [Pope et al. *Org Synth* **57** 45 1977, Keller et al. *Org Synth* **63** 160 1985.] It is a useful reagent for easy introduction of the *N*-Boc protecting group in amines [Iwanowicz et al. *Synth Commun* **23** 1443 1993], amino acids, peptides and proteins [Keller et al. *Org Synth* **63** 160 1985], amides [Flynn et al. *J Org Chem* **48** 2424 1983, Grehn et al. *Angew Chem* **97** 519 1985], and *N*-Boc-ylation of sensitive compounds in non-aqueous media [Kemp & Carey *J Org Chem* **54** 3640 1989]. Boc protection of alcohols has been achieved *via* Lewis acid catalysis [Bartoli et al. *Synlett* 2104 2006], and Boc reacted with 4-carboxyphenylhydrazine to form the *N,N,N'*-*tris*-Boc protected acid which was converted to 4-Fmoc(9-fluorenylmethoxycarbonyl ester)phenyl-hydrazine which is used as a chromophoric reagent to quantify aldehydes attached to a solid-phase since the hydrazine NH<sub>2</sub> group forms a coloured Schiff's base with the aldehyde [Shannon & Barany *J Org Chem* **69** 4586 2004]. [See the many volumes of Fieser and Fieser *Reagents for Organic Synthesis*, J Wiley & Sons, for more applications.]

***N,N*-Dibutyl formamide** [761-65-9] **M 157.3, b 63°/0.1mm, 118-120°/15mm, 244-246°/760mm, d<sub>4</sub><sup>20</sup> 0.878, n<sub>D</sub><sup>20</sup> 1.445.** Purify the amide by fractional distillation [Mandel & Hill *J Am Chem Soc* **76** 3981 1954]. [*Beilstein* **4** IV 565.]

**Di-*tert*-butyl peroxide (*tert*-butyl peroxide)** [110-05-4] **M 146.2, d 0.794, n<sub>D</sub><sup>20</sup> 1.389.** Wash the peroxide with aqueous AgNO<sub>3</sub> to remove olefinic impurities, water and dry (MgSO<sub>4</sub>). Free it from *tert*-butyl hydroperoxide by passage through an alumina column [Jackson et al. *J Am Chem Soc* **107** 208 1985], and if necessary two high vacuum distillations from room temperature to a liquid-air trap [Offenbach & Tobolsky *J Am Chem Soc* **79** 278 1957]. [Beilstein **1** IV 1619.] *The necessary protection from EXPLOSION should be used.*

**Di-*n*-butyl sulfide** [544-40-1] **M 146.3, α-form b 182°, β-form b 190-230°(dec).** Wash the sulfide with aqueous 5% NaOH, then water. Dry with CaCl<sub>2</sub> and distil it from sodium. [Beilstein **1** IV 1559.]

**Di-*n*-butyl sulfone** [598-04-9] **M 162.3, m 43.5°.** Purify it by zone melting. It crystallises from petroleum ether (**m** 44-44.5°), CHCl<sub>3</sub> (**m** 44°), and EtOH (**m** 45°). [Beilstein **1** H 371, **1** II 400, **1** III 1524, **1** IV 1561.]

***N,N'*-Di-*tert*-butylthiourea** [4041-95-6] **M 188.3, m 174-175°(evacuated capillary).** Recrystallise it from H<sub>2</sub>O [Bortnick et al. *J Am Chem Soc* **78** 4358 1956]. [Beilstein **4** IV 585 for *N*-butylthiourea.]

**Dichloroacetic acid** [79-43-6] **M 128.9, m 13.5°, b 95.0-95.5°/17-18mm, d<sub>4</sub><sup>20</sup> 1.563, n<sub>D</sub><sup>20</sup> 1.466, pK<sup>25</sup> 1.35.** Crystallise this strong acid from \*benzene or petroleum ether. Dry it with MgSO<sub>4</sub> and fractionally distil it. [Bernasconi et al. *J Am Chem Soc* **107** 3612 1985, Beilstein **2** IV 498.]

***sym*-Dichloroacetone (1,3-dichloropropan-2-one)** [534-07-6] **M 127.0, m 41-43°, 45°, b 86-88°/12mm, 75-77°/22mm, 172-172.5°/760mm, 170-175°/760mm, d<sub>4</sub><sup>20</sup> 1.383.** Crystallise it from CCl<sub>4</sub>, CHCl<sub>3</sub> or \*benzene and/or distil it under vacuum [Conant & Quayle *Org Synth Coll Vol I* 211 1941, Hall & Sirel *J Am Chem Soc* **74** 836 1952]. It is dimorphic [Daasch & Kagarise *J Am Chem Soc* **77** 6156 1955]. The *oxime* has **m** 130-131°, **b** 106°/25mm [*Arzneimittel-Forsch* **8** 638 1958]. [Beilstein **1** IV 3219.]

**Dichloroacetonitrile** [3018-12-0] **M 110.0, b 110-112°, d<sub>4</sub><sup>20</sup> 1.369, n<sub>D</sub><sup>20</sup> 1.440.** Purify the nitrile by distillation or by gas chromatography. [Beilstein **2** IV 506.] **FLAMMABLE.**

**2,5-Dichlorobenzoic acid** [50-79-3] **M 191.0, m 154°, b 301°/760mm, pK<sup>25</sup> 2.47.** Recrystallise the acid from water. [Beilstein **9** IV 1005.]

**2,3-Dichloro-1,3-butadiene** [1653-19-6] **M 123.0, b 41-43°/85mm, 98°/760mm.** Crystallise it from pentane to constant melting point of -40°. A mixture of *meso* and *d,l* forms is separated by gas chromatography on an 8m stainless steel column (8mm i.d.) with 20% DEGS (diethyleneglycolsilyl chloride) on Chromosorb W (60-80 mesh) at 60° and 80ml He/min. Distil it under vacuum. [Su & Ache *J Phys Chem* **80** 659 1976.]

**1,2-Dichloro-1,2-difluoroethane** [431-06-1] **M 134.9, b 59°, n<sub>D</sub><sup>20</sup> 1.376.** Purify it by fractional distillation [Hazeldine *J Chem Soc* 4258 1952]. For purification of a diastereoisomeric mixture, with resolution into *meso* and *rac* forms, see Machulla and Stocklin [*J Phys Chem* **78** 658 1974].

**Dichlorodifluoromethane (Freon 12)** [75-71-8] **M 120.9, m -158°, b -29.8°/atm, 42.5°/10atm.** Pass the gas through saturated aqueous KOH then conc H<sub>2</sub>SO<sub>4</sub>, and a tower packed with activated copper on Kiesselguhr at 200° to remove CO<sub>2</sub> and O<sub>2</sub>. A trap cooled to -29° removes a trace of high boiling material. It is a non-flammable propellant.

**1,1-Dichloroethane (ethylidene dichloride)** [75-34-3] **M 99.0, b 57.3°, d<sub>4</sub><sup>15</sup> 1.18350, d<sub>4</sub><sup>20</sup> 1.177, n<sub>D</sub><sup>15</sup> 1.41975.** Shake it with conc H<sub>2</sub>SO<sub>4</sub> or aqueous KMnO<sub>4</sub>, then wash it with water, saturated aqueous NaHCO<sub>3</sub>, again with water, dry with K<sub>2</sub>CO<sub>3</sub> and distil it from CaH<sub>2</sub> or CaSO<sub>4</sub>. Store it over silica gel. [Beilstein **1** IV 130.]

**1,2-Dichloroethane** [107-06-2] **M 99.0, b 83.4°, d<sub>4</sub><sup>20</sup> 1.256, n<sub>D</sub><sup>15</sup> 1.44759.** It is usually prepared by chlorinating ethylene, so that likely impurities include higher chloro derivatives and other chloro compounds depending on the impurities originally present in the ethylene. It forms azeotropes with water, MeOH, EtOH, trichloroethylene, CCl<sub>4</sub> and isopropanol. Its azeotrope with water (containing 8.9% water, has **b** 77°) can be

used to remove gross amounts of water prior to final drying. As a preliminary purification step, it can be steam distilled, and the lower layer is treated as below.

Shake it with conc  $\text{H}_2\text{SO}_4$  (to remove alcohol added as an oxidation inhibitor), wash with water, then dilute KOH or aqueous  $\text{Na}_2\text{CO}_3$  and again with water. After an initial drying with  $\text{CaCl}_2$ ,  $\text{MgSO}_4$  or by distillation, it is refluxed with  $\text{P}_2\text{O}_5$ ,  $\text{CaSO}_4$  or  $\text{CaH}_2$  and fractionally redistilled. Carbonyl-containing impurities can be removed as described for chloroform. [Beilstein 1 IV 131.]

**1,2-Dichloroethylene** [*cis* + *trans* 540-59-0] M 96.9, b 60° (*cis*),  $d_4^{20}$  1.284, b 48° (*trans*),  $d_4^{20}$  1.257. Shake it successively with conc  $\text{H}_2\text{SO}_4$ , water, aqueous  $\text{NaHCO}_3$  and water. Dry it with  $\text{MgSO}_4$  and fractionally distil it through an efficient column to separate the *cis*- and *trans*-isomers. [Beilstein 1 IV 707-709.]

***cis*-1,2-Dichloroethylene** [156-59-2] M 96.9, b 60.4°,  $d_4^{20}$  1.2830,  $n_D^{15}$  1.44903,  $n_D^{20}$  1.4495. Purify it by careful fractional distillation, followed by passage through neutral activated alumina. Also by shaking with mercury, drying with  $\text{K}_2\text{CO}_3$  and distilling from  $\text{CaSO}_4$ . Stabilise it with 0.02% of 2,6-di-*tert*-butyl-*p*-cresol. [Beilstein 1 IV 707.]

***trans*-1,2-Dichloroethylene** [156-60-5] M 96.9, b 47.7°,  $n_D^{15}$  1.45189,  $d_4^{20}$  1.2551,  $n_D^{20}$  1.4462. Dry it with  $\text{MgSO}_4$ , and fractionally distil it under  $\text{CO}_2$ . Fractional crystallisation at low temperatures has also been used. [Beilstein 1 IV 709.]

**2,3-Dichloromaleic anhydride** [1122-17-4] M 167.0, m 105-115°, 120°, 121-121.5°. Purify the anhydride by sublimation *in vacuo* [Katakis et al. *J Chem Soc, Dalton Trans* 1491 1986]. It has also been purified by Soxhlet extraction with hexane, recrystallisation from  $\text{CHCl}_3$  and by sublimation. [MS, Relles *J Org Chem* 37 3630 1972]. [Beilstein 17/11 V 63.]

**Dichloromethane (methylene dichloride)** [75-09-2] M 84.9, b 40.0°,  $d_4^{20}$  1.325,  $n_D^{20}$  1.42456,  $n_D^{25}$  1.4201. Shake it with portions of conc  $\text{H}_2\text{SO}_4$  until the acid layer remains colourless, then wash with water, aqueous 5%  $\text{Na}_2\text{CO}_3$ ,  $\text{NaHCO}_3$  or NaOH, then water again. Pre-dry with  $\text{CaCl}_2$ , and distil it from  $\text{CaSO}_4$ ,  $\text{CaH}_2$  or  $\text{P}_2\text{O}_5$ . Store it away from bright light in a brown bottle with Linde type 4A molecular sieves, in an atmosphere of dry  $\text{N}_2$ . Other purification steps include washing with aqueous  $\text{Na}_2\text{S}_2\text{O}_3$ , passage through a column of silica gel, and removal of carbonyl-containing impurities as described under **Chloroform**. It has also been purified by treatment with basic alumina, distillation, and stored over molecular sieves under nitrogen [Puchot et al. *J Am Chem Soc* 108 2353 1986].

Dichloromethane from Japanese sources contained MeOH as stabiliser which is not removed by distillation. It can, however, be removed by standing over activated 3A Molecular Sieves (note that 4A Sieves cause the development of pressure in bottles), passed through activated  $\text{Al}_2\text{O}_3$  and distilled [Gao et al. *J Am Chem Soc* 109 5771 1987]. It has been fractionated through a platinum spinning band column, degassed, and distilled onto degassed molecular sieves Linde 4A (heated under high vacuum at over 450° until the pressure readings reached the low values of  $10^{-6}$  mm, ~1-2 hours). Stabilise it with 0.02% of 2,6-di-*tert*-butyl-*p*-cresol [Mohammad & Kosower *J Am Chem Soc* 93 2713 1971]. [Beilstein 1 IV 35.]

**Rapid purification:** Reflux over  $\text{CaH}_2$  (5% w/v) and distil it. Store it over 4A molecular sieves.

**1,2-Dichloropropane** [78-87-5] M 113.0, b 95.9-96.2°,  $d_4^{20}$  1.158,  $n_D^{20}$  1.439. Distil the propane from  $\text{CaH}_2$ . It has a limited shelf life. [Beilstein 1 IV 195.]

**2,2-Dichloropropane** [594-20-7] M 113.0, b 69.3°,  $d_4^{20}$  1.090,  $n_D^{20}$  1.415. Wash it with aqueous  $\text{Na}_2\text{CO}_3$  solution, then distilled water, dry it over  $\text{CaCl}_2$  and fractionally distil it. [Beilstein 1 IV 196.]

**Di-*n*-decylamine** [1120-49-6] M 297.6, m 34°. b 153°/1mm, 359°/760mm,  $\text{pK}_{\text{Est}} \sim 11.0$ . Dissolve the amine in \*benzene and precipitate it as its bisulfate salt by shaking with 4M  $\text{H}_2\text{SO}_4$ . Filter, wash with \*benzene, separate by centrifugation, then the free base is obtained by treating with NaOH [McDowell & Allen *J Phys Chem* 65 1358 1961]. It is a strong base; store away from  $\text{CO}_2$ . [Beilstein 4 IV 780.]

**Didodecylamine** [3007-31-6] M 353.7, m 51.8°, b 263-265°/27mm,  $d_4^{25}$  0.806,  $\text{pK}^{25}$  11.00. Crystallise the

amine from EtOH/\*C<sub>6</sub>H<sub>6</sub> under N<sub>2</sub>, and store away from CO<sub>2</sub>. It provides two crystalline forms: an α-form with *m* 44.4° and a β-form with *m* 51.8°. The *hydrochloride* has *m* 207-208°(dec. from isoProOH), the *hydroiodide* has *m* 23.8-234°(dec. sealed capillary), and the *nitrate* has *m* 125.4-125.2°(dec. sealed capillary) when crystallised from MeOH/Me<sub>2</sub>CO. [Hoerr et al. *J Am Chem Soc* **65** 328 1943, Hoerr & Harwood *J Org Chem* **16** 779 1951, *Beilstein* **4** III 412, **4** IV 801.]

**Didodecyldimethylammonium bromide** [3282-73-3] *M* 463.6, *m* 157-162°. Recrystallise the salt from acetone, acetone/ether mixture, then from ethyl acetate, wash with ether and dry it in a vacuum oven at 60° [Chen et al. *J Phys Chem* **88** 1631 1984, Rupert et al. *J Am Chem Soc* **107** 2628 1985, Halpern et al. *J Am Chem Soc* **108** 3920 1986, Allen et al. *J Phys Chem* **91** 2320 1987]. [*Beilstein* **4** IV 801.]

**Diethanolamine (2,2'-iminodiethanol)** [111-42-2] *M* 105.1, *m* 28°, *b* 154-155°/10mm, 270°/760mm *pK*<sup>25</sup> 8.88. Fractionally distil the amine twice, then fractionally crystallise it from its melt. Its solubility in H<sub>2</sub>O is 10% at 20°. It absorbs CO<sub>2</sub> and HO<sub>2</sub> from the atmosphere. [Perrin & Dempsey *Buffers for pH and Metal Ion Control* Chapman & Hall, London 1974, *Beilstein* **4** H 283, **4** II 729, **4** III 689, **4** IV 1514.]

*N,N*-**Diethylacetamide** [685-91-6] *M* 157.2, *b* 86-88°, *d*<sub>4</sub><sup>20</sup> 0.994, *n*<sub>D</sub><sup>20</sup> 1.474. Dissolve the amide in cyclohexane, shake with anhydrous BaO and then filter. The procedure is repeated three times, and the cyclohexane is distilled off at atmospheric pressure. The crude amide is also fractionally distilled three times from anhydrous BaO. [*Beilstein* **4** III 349.]

**Diethyl acetamidomalonate** [1068-90-2] *M* 217.2, *m* 96°. Crystallise the ester from \*benzene petroleum ether. [*Beilstein* **4** III 2993.]

**Diethyl acetylenedicarboxylate** [762-21-0] *M* 170.2, *b* 60-62°/0.3mm, 107-110°/11mm, 118-120°/20mm, *d*<sub>4</sub><sup>20</sup> 1.0735, *n*<sub>D</sub><sup>20</sup> 1.4428. Dissolve the ester in \*C<sub>6</sub>H<sub>6</sub>, wash it with NaHCO<sub>3</sub>, H<sub>2</sub>O, dry over Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil it in a vacuum [IR: Walton & Hughes *J Am Chem Soc* **79** 3985 1957, Truce & Kruse *J Am Chem Soc* **81** 5372 1959]. [*Beilstein* **2** H 803.]

**Diethylamine** [109-89-7] *M* 73.1, *b* 55.5°, *d*<sub>4</sub><sup>20</sup> 0.707, *n* 1.38637, *pK*<sup>15</sup> 11.38. Dry diethylamine with LiAlH<sub>4</sub> or KOH pellets. Reflux it with, and distil it from, BaO or KOH. Convert it to the *p*-toluenesulfonamide and crystallise it to constant melting point from dry petroleum ether (*b* 90-120°), then hydrolyse it with HCl; excess NaOH is added, and the amine is passed through a column of activated alumina. Redistil the amine and dry it with activated alumina before use [Swift *J Am Chem Soc* **64** 115 1942]. [*Beilstein* **4** III 313.]  
§ A polystyrene diethylaminomethyl supported version is commercially available.

**Diethylamine hydrochloride** [660-68-4] *M* 109.6, *m* 223.5°, 226-229°. Crystallise salt from absolute EtOH. Also recrystallise it from dichloroethane/MeOH. *Hygroscopic*. [*Beilstein* **4** III 113.]

**Diethyl azodicarboxylate (DEAD)** [1972-28-7] *M* 174.2, *b* 104.5°/12mm, 211-213°/atm, *d*<sub>4</sub><sup>20</sup> 1.110, *n*<sub>D</sub><sup>20</sup> 1.420. Dissolve DEAD in toluene, wash it with 10% NaHCO<sub>3</sub> till neutral (may require several washes if too much hydrolysis had occurred: check IR for OH bands), then it wash with H<sub>2</sub>O (2x), dry it over Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate the toluene and distil it through a short Vigreux column at as high a vacuum as possible. The main portion boils at 107-111°/15mm. *Since it is likely to explode, use an oil bath for heating the still and all operations should be carried out behind an adequate shield with head protection*. [Rabjohn *Org Synth Coll Vol III* 375 1955, see Kauer *Org Synth Coll Vol IV* 412 1963]. [*Beilstein* **3** III 233.] It is commercially available as a 40% solution in toluene. This reagent is useful in the Mitsunobu reaction [Mitsunobu *Synthesis* **1** 1981, Gennari et al. *J Am Chem Soc* **108** 6394 1986, Evans et al. *J Am Chem Soc* **108** 6394 1986, Hughes *Org React* **42** 335 1992, Dodge et al. *Org Synth* **73** 110 1996, Hughes *Org Prep Proc Int* **28** 127 1996, Ferguson & Marcelle *J Am Chem Soc* **128** 4576 2006; see also di-*tert*-butyl azodicarboxylate above and **DIAD** below].

§ A polystyrene supported DEAD version is commercially available with a loading of ~1.2mmol/g.

**Diethyl bromomalonate** [685-87-0] *M* 239.1, *b* 116-118°/10mm, 122-123°/20mm, *d*<sub>4</sub><sup>20</sup> 1.420, *n*<sub>D</sub><sup>20</sup> 1.4507. Purify the ester by fractional distillation in a vacuum. Its IR (film) has *v*<sub>max</sub> 1800 and 1700cm<sup>-1</sup> [Abramovitch

*Can J Chem* **37** 1146 1959, Bretschneider & Karpitschka *Monatsh Chem* **84** 1091 1053]. [*Beilstein* **2** IV 1904.]

**Diethyl tert-butylmalonate** [759-24-0] **M** 216.3, **b** 40-42°/0.03, 102-104°/11mm, 109.5-110.5°/17mm, 205-210°/760mm,  $d_4^{20}$  0.980,  $n_D^{20}$  1.425. Dissolve it in Et<sub>2</sub>O, wash with aqueous NaHCO<sub>3</sub>, H<sub>2</sub>O, dry (MgSO<sub>4</sub>), filter, evaporate and distil the residue. Identify by hydrolysis to the acid and determine the neutralisation equivalent (theor: 80.0). The acid has **m** 155-157° effervescence [Hauser et al. *J Am Chem Soc* **64** 2715 1942, Bush & Beauchamp *J Am Chem Soc* **75** 2949 1953]. [*Beilstein* **2** IV 2027.]

**Diethyl carbonate** [105-58-8] **M** 118.1, **b** 124-125°, 126.8°,  $d_4^{20}$  0.975,  $n_D^{25}$  1.38287. Wash the ester (100ml) with an aqueous 10% Na<sub>2</sub>CO<sub>3</sub> (20ml) solution, saturated CaCl<sub>2</sub> (20ml), then water (30ml). After drying by standing over solid CaCl<sub>2</sub> for 1 hour (note that prolonged contact should be avoided because slow combination with CaCl<sub>2</sub> occurs), it should be fractionally distilled. Also dry it over MgSO<sub>4</sub> and distil it. [*Beilstein* **3** H 5, **3** I 4, **3** II 4, **3** III 5, **3** IV 5.]

**Diethyl disulfide** [110-81-6] **M** 122.3, **b** 86.2°/87mm, 154-155°/atm,  $d_4^{20}$  0.993,  $n_D^{20}$  1.506. Dry the disulfide over silica gel or MgSO<sub>4</sub> and distil it under reduced pressure (optionally from CaCl<sub>2</sub>). [*Beilstein* **1** H 347, **1** I 173, **1** II 345, **1** III 1377, **1** IV 1379.]

**Diethylene glycol** [111-46-6] **M** 106.1, **f** -10.5°, **b** 244.3°,  $d_4^{20}$  1.118,  $n_D^{15}$  1.4490,  $n_D^{20}$  1.4475. Fractionally distil it in a vacuum (**b** 133°/14mm, 2.5cm x 1.3m heli-grid column), then recrystallise it by partial freezing. [Feldman et al. *J Am Chem Soc* **73** 4341 1951, *Beilstein* **1** III 2090, **1** IV 2390.]

**Diethylene glycol diethyl ether** [112-36-7] **M** 162.2, **b** 76°/32mm, 85-86°/10mm, 188.2-188.3°/751mm,  $d_4^{20}$  0.910,  $n_D^{20}$  1.412. Dry the ether with MgSO<sub>4</sub>, then CaH<sub>2</sub> or LiAlH<sub>4</sub>, under N<sub>2</sub>. If sodium is used, the ether should be redistilled alone to remove any products which may be formed by the action of sodium on the ether. As a preliminary purification, the crude ether (2L) can be refluxed for 12 hours with 25ml of conc HCl in 200ml of water, under reduced pressure, with slow passage of N<sub>2</sub> to remove aldehydes and other volatile substances. After cooling, add sufficient solid KOH pellets (slowly and with shaking until no more dissolves) to give two liquid phases. The upper of these is decanted, dried with fresh KOH pellets, decanted, then refluxed over, and distilled from sodium. It can be passed through (alkaline) alumina prior to purification. [*Beilstein* **1** IV 2394.]

**Diethylene glycol ditosylate** [7460-82-4] **M** 414.5, **m** 86-87°, 87-88°, 88-89°. Purify the ester by recrystallisation from Me<sub>2</sub>CO and dry it in a vacuum. [*Beilstein* **11** III 225.]

**Diethylene glycol mono-*n*-butyl ether (butyl carbitol)** [112-34-5] **M** 162.2, **b** 69-70°/0.3mm, 230.5°/760mm,  $d_4^{20}$  0.967,  $n_D^{20}$  1.4286. Dry the ether with anhydrous K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, filter and fractionally distil it. Peroxides can be removed by refluxing with stannous chloride or a mixture of FeSO<sub>4</sub> and KHSO<sub>4</sub> (or, less completely, by filtration under slight pressure through a column of activated alumina). [*Beilstein* **1** IV 2394.]

**Diethylene glycol monoethyl ether** [111-90-0] **M** 134.2, **b** 201.9°,  $d_4^{20}$  0.999,  $n_D^{20}$  1.4273,  $n_D^{25}$  1.4254. Ethylene glycol can be removed by extracting 250g in 750ml of benzene with 5ml portions of water, allowing for phase separation, until successive aqueous portions show the same volume increase. Dry, and free from peroxides, as described for diethylene glycol mono-*n*-butyl ether. [*Beilstein* **1** IV 2393.]

**Diethylene glycol monomethyl ether** [111-77-3] **M** 120.2, **b** 194°,  $d_4^{20}$  1.010,  $n_D^{20}$  1.423. Purify it as for diethylene glycol mono-*n*-butyl ether. [*Beilstein* **1** IV 2392.]

**Diethylenetriaminepenta-acetic acid (DTPA, DEPTAPAC)** [67-43-6] **M** 393.4, **m** 219-220°,  $pK_1^{25}$  1.79,  $pK_2^{25}$  2.56,  $pK_3^{25}$  4.42,  $pK_4^{25}$  8.76,  $pK_5^{25}$  10.42. Crystallise DTPA from water. Dry it under vacuum or at 110°. [Bielski & Thomas *J Am Chem Soc* **109** 7761 1987, NMR: Wenzel et al. *Anal Chem* **54** 615 1982, *Beilstein* **4** IV 2454.]

**Diethyl ether (ethyl ether)** [60-29-7] **M** 74.1, **b** 34.6°/760mm,  $d_4^{20}$  0.714,  $n_D^{15}$  1.3555,  $n_D^{20}$  1.35272. Usual

impurities are water, EtOH, diethyl peroxide (which is explosive when concentrated), and aldehydes. Peroxides [detected by liberation of iodine from weakly acid (HCl) solutions of KI, or by the blue colour in the ether layer when 1mg of  $\text{Na}_2\text{Cr}_2\text{O}_7$  and 1 drop of dilute  $\text{H}_2\text{SO}_4$  in 1ml of water is shaken with 10ml of ether] can be removed in several different ways. The simplest method is to pass dry ether through a column of activated alumina (80g  $\text{Al}_2\text{O}_3/700\text{ml}$  of ether). More commonly, 1L of ether is shaken repeatedly with 5-10ml of a solution comprising 6.0g of ferrous sulfate and 6ml of conc  $\text{H}_2\text{SO}_4$  in 110ml of water. Aqueous 10%  $\text{Na}_2\text{SO}_3$  or stannous chloride can also be used. The ether is then washed with water, dried for 24 hours with  $\text{CaCl}_2$ , filtered and dried further by adding sodium wire until it remains bright. The ether is stored in a dark cool place, until distilled from sodium before use. Peroxides can also be removed by wetting the ether with a little water, then adding excess  $\text{LiAlH}_4$  or  $\text{CaH}_2$  and leaving to stand for several hours. (This also dried the ether.)

Werner [*Analyst* **58** 335 1933] removed peroxides and aldehydes by adding 8g  $\text{AgNO}_3$  in 60ml of water to 1L of ether, then 100ml of 4% NaOH and shaking for 6 minutes. Fierz-David [*Chimia* **1** 246 1947] shook 1L of ether with 10g of a zinc-copper couple. (This reagent is prepared by suspending zinc dust in 50ml of hot water, adding 5ml of 2M HCl and decanting after 20 seconds, washing twice with water, covering with 50ml of water and 5ml of 5% cuprous sulfate with swirling. The liquid is decanted and discarded, and the residue is washed three times with 20ml of ethanol and twice with 20ml of diethyl ether).

Aldehydes can be removed from diethyl ether by distillation from hydrazine hydrogen sulfate, phenylhydrazine or thiosemicarbazide. Peroxides and oxidisable impurities have also been removed by shaking with strongly alkaline saturated- $\text{KMnO}_4$  (with which the ether was left to stand in contact for 24 hours), followed by washing with water, conc  $\text{H}_2\text{SO}_4$ , water again, then drying ( $\text{CaCl}_2$ ) and distillation from sodium, or sodium containing benzophenone to form the ketyl. Other purification procedures include distillation from sodium triphenylmethide or butyl magnesium bromide, and drying with solid NaOH or  $\text{P}_2\text{O}_5$ . [*Beilstein* **1** IV 1314.]

**Rapid purification:** Same as for 1,4-dioxane.

**Diethyl ethoxymethylene malonate** [87-13-8] **M 216.2, b 014°/0.2mm, 109°/0.5mm, 279-283°/atm,  $d_4^{20}$  1.079,  $n_D^{20}$  1.4623.** Likely impurity is diethyl diethoxymethylene malonate which is difficult to separate from diethyl ethoxymethylene malonate by distillation, and it is necessary to follow the course of the distillation by the change in refractive index instead of boiling point. After a low boiling fraction is collected, there is obtained an intermediate fraction ( $n_D^{20}$  1.414—1.458), the size of which depends on the amount of the diethoxymethylene compound. This fraction is fractionated through a 5inch Vigreux column at low pressure, and avoiding interruption in heating. Fraction **b** 108-110°/0.25mm is *ca* 10° lower than the submitters' fraction (**b** 97.2°/0.25mm,  $n_D^{20}$  1.4612—1.4623) [*Org Synth Coll Vol III* 395 1955, Fuson et al. *J Org Chem* **11** 197 1946, Duff & Kendal *J Chem Soc* 893 1948]. [*Beilstein* **3** IV 1192.]

***N,N'*-Dimethylethylenediamine [1,2-bis(methylamino)ethane]** [110-70-3] **M 88.2, b 110-112°/750mm, 119°/760mm,  $d_4^{20}$  0.819,  $n_D^{20}$  1.431, (pK<sup>25</sup> 7.01 and 9.88).** This strong base has been prepared in various ways including hydrolysis of *N,N'*-di(benzylsulfonyl)-*N,N'*-dimethylethylenediamine (**m** 217-219°, from AcOH) and concentrated HCl (120-130°) [Johnson & Bailey *J Am Chem Soc* **38** 2135 1916], the reaction of 1,2-dibromoethane with 21-33% aqueous  $\text{MeNH}_2$  in EtOH (reflux for 2 hours, 50% yield) [Kermack & Wight *J Chem Soc* 1421 1935, Woodburn & O'Gee *J Org Chem* **17** 1235 1952], hydrolysis of *N,N'*-di(*p*-toluenesulfonyl)-*N,N'*-dimethylethylenediamine (1mol, **m** 164°, from AcOH) with  $\text{H}_2\text{SO}_4$  (8.2mol) and  $\text{H}_2\text{O}$  (9mol) at 140-145°/7 hours [Boon *J Chem Soc* 307 1947], LAH reduction in THF of 3-benzenesulfonyloxy-5,6-dihydrouracil (**m** 175-176°, from iso-PrOH) and isolated as the *dibenzoyl derivative* (41%, **m** 177-178°, from  $^*\text{C}_6\text{H}_6$ ) [Bauer *J Am Chem Soc* **78** 1945 1956], and by catalytic hydrogenation of *N*-benzyl-*N,N'*-dimethylethylenediamine (**b** 73-74°/0.1mm) with 10% Pd/C [Jucker & Rissi *Helv Chim Acta* **45** 2383 1962]. General isolation and purification procedures involve steam distillation of the product diamine, acidifying the distillate with HCl, evaporating to dryness, treating the residual salt with 10% excess of cold 32% aqueous NaOH, collecting the organic layer (**care**: very caustic solution), drying it over solid KOH, and distilling from it. Final distillation of the dry base over Na has been reported (see also [107-15-3]). The diamine is a strong base and readily absorbs  $\text{CO}_2$  and  $\text{H}_2\text{O}$  from the atmosphere—store it in a dark stoppered bottle, wax the stopper if it should be stored for long periods. The *dihydrochloride* [5752-40-9] has **m** 235-236°(dec), and the *picrate* has **m** 160° (from EtOH and  $\text{Me}_2\text{CO}$  in rectangular plates). It forms complexes with Cu, Ni, and Pt among other metals. Its FT-IR (neat) has  $\nu_{\text{max}}$  at

3285.5, 2788.2, 1447.2, 1346.5, 1251.0, 1106.5, 1040.5, 876.9, 763.7  $\text{cm}^{-1}$ . [For pK and metal complexes see Gastafson & Martell *J Am Chem Soc* **81** 525 1959, *Beilstein* **4** H 250, **4** I 415, **4** II 689, **4** III 512, **4** IV 1171.]

*N,N'*-Diethylformamide [617-84-5] **M** 101.2, **b** 29°/0.5mm, 61-63°/10mm, 178.3-178.5°/760mm, **d**<sub>4</sub><sup>20</sup> 0.906, **n**<sub>D</sub><sup>25</sup> 1.4313. Distil it under reduced pressure, then at atmospheric pressure [Wintcler et al. *Helv Chim Acta* **37** 2370 1954, NMR: Hoffmann *Z Anal Chem* **170** 177 1959]. [*Beilstein* **4** IV 346.]

**Diethyl fumarate** [623-91-6] **M** 172.2, **b** 218°, **d** 1.052, **n** 1.441. Wash the fumarate with aqueous 5%  $\text{Na}_2\text{CO}_3$ , then with saturated  $\text{CaCl}_2$  solution, dry with  $\text{CaCl}_2$  and distil it. [*Beilstein* **2** IV 2207.] Note that **dimethyl fumarate** [624-49-7] **M** 144.1, **m** 102°, **b** 192-193°/atm is a solid. [*Beilstein* **2** IV 2205.]

*N,N*-Diethyl-1,1,2,3,3,3-hexafluoropropylamine (**Ishikawa's Reagent, perfluoropropyldiethylamine, PPDA**) [309-88-6] **M** 223.2, **b** 56-57°/58mm **d**<sub>4</sub><sup>25</sup> 1.230, **n**<sub>D</sub><sup>20</sup> 1.3460. When this reagent is prepared from diethylamine (11g, 0.15mol) in dry  $\text{Et}_2\text{O}$  (30ml) in a glass pressure bomb cooled to  $-70^\circ$  ( $\text{Me}_2\text{CO}$ /Dry Ice), and liquefied *F*-propene (25.5g, 0.17mol, see[116-15-4]) is added, and the temperature of the sealed bomb is allowed to rise to  $\sim 25^\circ$ , and stirred at this temperature overnight, the reaction is complete. Crystalline  $\text{Et}_2\text{NH}\cdot\text{HF}$  is removed by filtration, the solvent ( $\text{Et}_2\text{O}$ ) is evaporated, and the residual oil is distilled in a vacuum to give a liquid (23.7g,  $\sim 72\%$ ) which boils at **56-57°/58mm**. This liquid is shown, by its  $^{19}\text{F}$  NMR signal intensities, to be a 3:1 mixture of PPDA and exclusively (*E*)-*N,N*-diethyl-pentafluoro-propenylamine (*trans*-dppa) with respectively  $\delta_{\text{F}}$  [PPDA] at -3.0 (1- $\text{CF}_3$ ), 6.0 (3- $\text{CF}_A$ ), 11.0 (3- $\text{CF}_B$ ), 131.0 (2-CF), and  $\delta_{\text{F}}$  [*trans*-dppa] at -12.0 (1- $\text{CF}_3$ ,  $^{1,3}J = 14\text{Hz}$ ,  $^{1,4}J = 23.5\text{Hz}$ ), 41.0 (3-CF,  $^{1,4}J = 23.5\text{Hz}$ ,  $^{1,3}J_{\text{trans}} = 117\text{Hz}$ ), 122 (2-CF,  $^{1,3}J = 14\text{Hz}$ ,  $^{1,3}J_{\text{trans}} = 117\text{Hz}$ ) upfield from external  $\text{CF}_3\text{CO}_2\text{H}$  with 3:1 relative intensities. The mixture is easy to handle, and is stable on storage at room temperature without discolouration or loss of activity for at least 6 months. It reacts with a variety of CHO groups to form the respective CHF derivatives, with carboxylic acids to form the acid fluorides [Takaoka, Iwakiri and Ishikawa *Bull Chem Soc Jpn* **11** 3377 1979, Watanabe, Fujita, Ushi and Kitazume *J Fluorine Chem* **31** 247 1986, Watanabe, Fujita, Sakamoto, Kuramochi and Kitazume *J Fluorine Chem* **36** 361 1987], and allylic alcohols to form  $\alpha$ -fluoro- $\alpha$ -(trifluoromethyl)- $\gamma,\delta$ -unsaturated amide precursors [Ogu, Akazome and Ogura *Tetrahedron Lett* **39** 305 1998]. The crude 3:1 reaction product has been used satisfactorily without distillation. The presence of 33% of *trans*-dppa is not a drawback because HF is liberated in these reaction of PPDA with OH, and the HF that is liberated adds on to the olefinic *trans*-dppa to give PPDA; so that as long as there is at least equal amounts of these two compounds the mixture behaves as if it were completely made up of DDPA.

When the above synthesis is carried out at  $\sim 5$ - $10^\circ$  while *F*-propene is bubbled through (completely absorbed within 2 hours), then stirred at  $\sim 25^\circ$  overnight, an oil is obtained (89% yield) with **51-53°/58mm** which is a 1:1 mixture (by  $^{19}\text{F}$  NMR) of PPDA and *trans*-dppa. If *n*- $\text{Bu}_2\text{NH}$ , or piperidine, replace  $\text{Et}_2\text{NH}$  in the reaction at  $-70^\circ$  in sealed bombs, (*n*- $\text{Bu}$ )<sub>2</sub>N- $\text{CF}_2\text{CHF}_2\text{CF}_3$  and (*E*)-(*n*- $\text{Bu}$ )<sub>2</sub>N- $\text{CF}=\text{CF}_2\text{CF}_3$  (**55-57°/5mm**, 2:1, 78% yield) or  $\text{C}_5\text{H}_{10}\text{N}-\text{CF}_2\text{CHF}_2\text{CF}_3$  only (**49-50°/7mm**, 80% yield) respectively are obtained [Takaoka, Iwakiri and Ishikawa *Bull Chem Soc Jpn* **11** 3377 1979].

**Diethyl ketone (3-pentanone)** [96-22-0] **M** 86.1, **b** 102.1°, **d**<sub>4</sub><sup>20</sup> 0.8099, **n**<sub>D</sub><sup>20</sup> 1.392. The ketone is dried with anhydrous  $\text{CaSO}_4$  or  $\text{CuSO}_4$ , and distil from  $\text{P}_2\text{O}_5$  under  $\text{N}_2$  or under reduced pressure. Further purification is by conversion to the semicarbazone (recrystallise to constant **m** 139°, from  $\text{EtOH}$ ) which, after drying *in vacuo* over  $\text{CaCl}_2$  and paraffin wax, is refluxed for 30 minutes with excess oxalic acid, then steam distilled and salted out with  $\text{K}_2\text{CO}_3$ . Dry with  $\text{Na}_2\text{SO}_4$  and distil [Cowan et al. *J Chem Soc* 171 1940]. [*Beilstein* **1** IV 3279.]

**Diethyl malonate** [105-53-3] **M** 160.2, **b** 92°/22mm, 198-199°/760mm, **d**<sub>4</sub><sup>20</sup> 1.056, **d**<sup>25</sup> 1.0507, **n**<sub>D</sub><sup>20</sup> 1.413. If too impure (IR, NMR) the ester (250g) is heated on a steam bath for 36 hours with absolute  $\text{EtOH}$  (125ml) and conc  $\text{H}_2\text{SO}_4$  (75ml), then fractionally distilled under reduced pressure. Otherwise fractionally distil it under reduced pressure and collect the steady boiling middle fraction. [*Beilstein* **2** IV 1881.]

**2,2-Diethyl-1,3-propanediol** [115-76-4] **M** 132.2, **m** 61.4-61.8°, **b** 130-133°/16mm, **n**<sub>D</sub><sup>20</sup> 1.4574. Crystallise the diol from petroleum ether (b 65-70°). **IRRITANT**. [McKusick *J Am Chem. Soc* **70** 1982 1948, *Beilstein* **1**

III 2217, 1 IV 2589.]

**Diethyl pyrocarbonate (DEP)** [1609-47-8] **M 162.1, b 38-40°/12mm, 160-163°/atm,  $d_4^{20}$  1.119,  $n_D^{20}$  1.398.** Dissolve the ester in Et<sub>2</sub>O, wash it with dilute HCl, H<sub>2</sub>O, dry over Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil the residue first *in vacuo* then at atmospheric pressure. It is soluble in alcohols, esters, ketones and hydrocarbon solvents. A 50% w/w solution is usually prepared for general use. **Treat with great CAUTION as DEP irritates the eyes, mucous membranes and skin.** [Boehm & Mehta *Chem Ber* 71 1797 1938, Thoma & Rinke *Justus Liebigs Ann Chem* 624 30 1959, *Beilstein* 3 IV 18.]

**Diethyl succinate** [123-25-1] **M 174.2, b 105°/15mm, 218°/atm,  $d_4^{20}$  1.047,  $n_D^{20}$  1.4199.** Dry the succinate with MgSO<sub>4</sub>, and distil it at 15mm pressure. [*Beilstein* 2 IV 1914.]

**Diethyl sulfate** [64-67-5] **M 154.2, b 96°/15mm, 118°/40mm,  $d_4^{20}$  1.177,  $n_D^{20}$  1.399.** Wash the ester with aqueous 3% Na<sub>2</sub>CO<sub>3</sub> (to remove acidic material), then distilled water, dry (CaCl<sub>2</sub>), filter and distil it in a vacuum. *It is an ethylating agent and blisters the skin.* [*Beilstein* 1 IV 1236.]

**Diethyl sulfide** [352-93-2] **M 90.2, m 0°/15mm, b 90.1°/760mm,  $d_4^{20}$  0.837,  $n_D^{20}$  1.443.** Wash the sulfide with aqueous 5% NaOH, then water, dry with CaCl<sub>2</sub> and distil it from sodium. It can also be dried with MgSO<sub>4</sub> or silica gel. *Alternative*, purification is *via* the Hg(II) chloride complex [(Et)<sub>2</sub>S.2HgCl<sub>2</sub>] (see dimethyl sulfide). [*Beilstein* 1 IV 1394.]

**Diethyl (–)-D- (from the non-natural) [13811-71-7] and (+)-L- (from the natural acid) [89-91-2] tartrate** **M 206.2, m 17°, b 80°/0.5mm, 162°/19mm, 278-282°/atm,  $d_4^{20}$  1.204,  $n_D^{20}$  1.4476,  $[\alpha]_D^{20}$  (+) and (+) 26.5° (c 1, H<sub>2</sub>O) and (–) and (+) 8.5° (neat),  $[\alpha]_{546}^{20}$  (+) and (+) 30° (c 1, H<sub>2</sub>O).** Distil the esters under high vacuum and store them under vacuum or in an inert atmosphere in a desiccator in round bottomed flasks equipped with a vacuum stopcock. They have also been distilled by Kügelrohr distillation and/or by 'wiped-film' molecular distillation. They are slightly soluble in H<sub>2</sub>O but miscible with EtOH and Et<sub>2</sub>O. [Gao et al. *J Am Chem Soc* 109 5770 (5771) 1987, IR: Pristera *Anal Chem* 25 844 1953, *Beilstein* 3 III 1025 for D(–), 3 IV 1232 for L(+).]

**sym-Diethylthiourea** [105-55-5] **M 132.2, m 76-77°. 77-79°.** Crystallise it from \*benzene. [*Beilstein* 4 H 118, 4 I 355, 4 II 610, 4 III 220, 4 IV 375.]

**Difluoroacetic acid** [381-73-7] **M 96.0, m –0.35°, b 67-70°/20mm, 134°/760mm,  $d_4^{20}$  1.530,  $n_D^{20}$  1.3428,  $pK_{25}^{25}$  1.28.** Purify the acid by distilling over P<sub>2</sub>O<sub>5</sub>. The *acid chloride* is a fuming liquid **b 25°/atm**, the *amide* has **b 108.6°/35mm, m 52°** (from \*C<sub>6</sub>H<sub>6</sub>), and the *anilide* has **b 90°/1mm, 114°/5mm, m 58°** [Henne & Pelley *J Am Chem Soc* 74 1426 1952, Coffman et al. *J Org Chem* 14 749 1949, NMR: Meyer et al. *J Am Chem Soc* 75 4567 1953,  $pK$ : Wegscheider *Z Phys Chem* 69 614 1909]. [*Beilstein* 2 IV 455.]

**Diglycolic acid (2-oxapentane-1,5-dioic acid)** [110-99-6] **M 134.1, m 148° (monohydrate),  $pK_1^{25}$  2.97,  $pK_2^{25}$  4.37.** Crystallise diglycolic acid from water. [*Beilstein* 3 IV 577.]

**Diglyme [bis(2-methoxyethyl) ether, diethylene glycol dimethyl ether]** [111-96-6] **M 134.2, b 62°/17mm, 75°/35mm, 160°/760mm,  $d_4^{20}$  0.917,  $n_D^{20}$  1.4087.** Dry diglyme with NaOH pellets or CaH<sub>2</sub>, then reflux with, and distil it (under reduced pressure) from Na, CaH<sub>2</sub>, LiAlH<sub>4</sub>, NaBH<sub>4</sub> or NaH. These operations are carried out under N<sub>2</sub>. The amine-like odour of diglyme has been removed by shaking with a weakly acidic ion-exchange resin (Amberlite IR-120) before drying and distilling. Addition of 0.01% NaBH<sub>4</sub> to the distillate inhibits peroxidation. Purify it also as for dioxane. It has been passed through a 12-in column of molecular sieves to remove water and peroxides. [*Beilstein* 1 IV 2393.]

**Dihydroxyfumaric (1,2-dihydroxybut-1-ene-1,2-dioic) acid dihydrate** [133-38-0] **M 184.1, m 155°(dec),  $pK_1^{25}$  1.57,  $pK_2^{25}$  3.36.** Crystallise the acid from water. [*Beilstein* 3 IV 1975.]

**1,2-Diiodoethane** [624-73-7] **M 281.9, m 81-84°, d 2.134.** Dissolve it in ether, wash it with saturated aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, dry it over MgSO<sub>4</sub>, and evaporate the ether *in vacuo* then distil it. Store it in the dark.

[Molander et al. *J Am Chem Soc* **109** 453 1987]. [*Beilstein* **1** IV 169.]

**Diiodomethane (methylene diiodide)** [75-11-6] **M 267.8, m 6.1°, b 66-70°/11-12mm, d<sub>4</sub><sup>20</sup> 3.325.** Fractionally distil it under reduced pressure, then fractionally crystallise it by partial freezing, and stabilise it with silver wool if necessary. It has also been purified by drying over CaCl<sub>2</sub> and fractionally distilling from Cu powder. Store it in the dark. [*Beilstein* **1** IV 97.]

**Diisopropanolamine** [110-97-4] **M 133.2, m 41-44°, d<sub>4</sub><sup>20</sup> 1.004, pK<sub>Est</sub> ~10.7.** Crystallise the amine repeatedly from dry diethyl ether. It is a strong base, store away from CO<sub>2</sub> and H<sub>2</sub>O. [*Beilstein* **4** III 761.]

**Diisopropylamine** [108-18-9] **M 101.2, b 83.5°/760mm, d<sub>4</sub><sup>20</sup> 0.720, n<sub>D</sub><sup>20</sup> 1.39236, pK<sup>25</sup> 11.20.** Distil the amine from NaOH, or reflux it three minutes over Na wire or NaH, and distil it into a dry receiver under N<sub>2</sub>. It is a strong base, store away from CO<sub>2</sub> and H<sub>2</sub>O. [*Beilstein* **4** H 154, **4** I 369, **4** II 630, **4** III 274, **4** IV 510.]  
§ A polystyrene supported version of diisopropylamine is commercially available.

**Diisopropylazodicarboxylate (DIAD)** [2446-83-5] **M 202.2, b 75°/0.2mm, d<sup>25</sup> 1.420, n<sub>D</sub><sup>20</sup> 1.420.** Purify the azo compound by distillation at as high a vacuum as possible. Since it is likely to explode, use an oil bath for heating the still, and all operations should be carried out behind an adequate shield. [Kauer *Org Synth Coll Vol IV* 412 1963, *Beilstein* **3** III 233]. This reagent is useful in the Mitsunobu reaction [Mitsunobu *Synthesis* **1** 1981, Gennari et al. *J Am Chem Soc* **108** 6394 1986, Evans et al. *J Am Chem Soc* **108** 6394 1986, Hughes *Org React* **42** 335 1992, Dodge et al. *Org Synth* **73** 110 1996, Hughes *Org Prep Proc Int* **28** 127 1996, Ferguson & Marcelle *J Am Chem Soc* **128** 4576 2006; see also di-*tert*-butyl azodicarboxylate and **DEAD** above].

**Diisopropylethylamine (Hünig's base)** [7087-68-5] **M 129.3, b 119°/731mm, 127°/760mm, d<sub>4</sub><sup>27</sup> 1.440, n<sub>D</sub><sup>25</sup> 1.4376, pK<sub>Est</sub> ~10.9.** Distil the amine from ninhydrin, then from KOH [Dryland & Sheppard, *J Chem Soc, Faraday Trans 1* 125 1986]. It is a strong base and should be stored in the absence of carbon dioxide. [Hünig & Kiessel *Chem Ber* **91** 380, 387 1958, Wotiz et al. *J Org Chem* **24** 1202 1959, *Beilstein* **4** IV 551.]

**Diisopropyl ketone (2,4-dimethyl-3-pentanone)** [565-80-0] **M 114.2, b 124°, d 0.801, n 1.400.** Dry the ketone with CaSO<sub>4</sub>, shake it with chromatographic alumina and fractionally distil it from P<sub>2</sub>O<sub>5</sub> under N<sub>2</sub>. [*Beilstein* **1** IV 3334.]

**Diketene** See 2-methylene-oxetan-2-one in "Heterocyclic Compounds", in this Chapter, and see ketene below.

**2,3-Dimercapto-1-propanol (BAL, British Anti-Lewisite)** [59-52-9] **M 124.2, b 82-84°/0.8mm, 120°/15mm, d<sub>4</sub><sup>20</sup> 1.239, n<sub>D</sub><sup>20</sup> 1.5732, pK<sub>1</sub><sup>25</sup> 8.62, pK<sub>2</sub><sup>25</sup> 10.75.** Precipitate BAL as the Hg mercaptide [see Bjöberg *Chem Ber* **75** 13 1942], regenerate with H<sub>2</sub>S, and distil it under a vacuum [Rosenblatt & Jean *Anal Chem* 951 1955]. *It is an antidote for heavy metal (As, Hg, Au, etc.) poisoning.* [*Beilstein* **1** IV 2770.]

**1,3-Dimercapto-2-propanol** [584-04-3] **M 124.2, b 68-69°/0.8mm, 82°/1.5mm, 94°/12mm, n<sub>D</sub><sup>23</sup> 1.5696.** Purify the dithiol as for 2,3-dimercapto-1-propanol above. [Johary & Owen *J Chem Soc* 1305 1955, *Beilstein* **1** IV 2773 or **2** IV 1102.]

**meso-2,3-Dimercaptosuccinic acid** [304-55-2] **M 182.2, m 191-192°(dec), 210°(dec), 210-211° (dec), pK<sub>1</sub><sup>25</sup> 2.71, pK<sub>2</sub><sup>25</sup> 3.48, pK<sub>3</sub><sup>25</sup> 8.89, pK<sub>4</sub><sup>25</sup> 10.75.** Purify the acid by dissolving it in NaOH and precipitating with dilute HCl, drying and recrystallising from MeOH. IR has  $\nu$  at 2544 (SH) and 1689 (CO<sub>2</sub>H) cm<sup>-1</sup>. The *bis-S-acetyl* derivative has **m** 183-185° (from EtOAc or Me<sub>2</sub>CO), and its *Me ester* has **m** 119-120° (from petroleum ether) [Gerecke et al. *Helv Chim Acta* **44** 957 1961, Owen & Sultanbawa *J Chem Soc* 3112 1949]. [*Beilstein* **3** III 1033.]

**1,2-Dimethoxyethane (DME, glycol dimethyl ether, glyme)** [110-71-4] **M 90.1, b 84°, d<sub>4</sub><sup>20</sup> 0.867, n 1.380.** Traces of water and acidic materials have been removed from it by refluxing with Na, K or CaH<sub>2</sub>, decanting and distilling from Na, K, CaH<sub>2</sub> or LiAlH<sub>4</sub>. The reaction has been speeded up by using vigorous high-speed stirring with molten potassium. For virtually complete elimination of water, 1,2-dimethoxyethane has been dried with

Na-K alloy until a characteristic blue colour is formed in the solvent at Dry-ice/cellosolve temperatures: the solvent is kept with the alloy until distilled for use [Ward *J Am Chem Soc* **83** 1296 1961]. Alternatively, glyme, refluxed with benzophenone and Na-K, is dry enough if, on distillation, it gives a blue colour of the ketyl immediately on addition to benzophenone and sodium [Ayscough & Wilson *J Chem Soc* 5412 1963]. It has also been purified by distillation under N<sub>2</sub> from sodium benzophenone ketyl (see above). [Beilstein **1** IV 2376.]

***N,N*-Dimethylacetamide (DMAc)** [127-19-5] **M 87.1, b 58.0-58.5°/11.4mm, 66-67°/15mm, 85-87°/33mm, 96°/80mm, 163-165°/760mm, d<sub>4</sub><sup>20</sup> 0.940, n<sub>D</sub><sup>20</sup> 1.4373.** Shake the amide with BaO for several days, reflux it over BaO for 1 hour, then fractionally distil it under reduced pressure. Store it over molecular sieves. It is a useful organic solvent for reactions, as it is soluble in many organic solvents and in H<sub>2</sub>O. Use it in a fume hood because over-exposure causes liver damage, jaundice, lethargy and can irritate skin. [Beilstein **4** IV 180.]

***β,β*-Dimethylacrylic acid (senecioic acid, 3-methyl-2-butenic acid)** [541-47-9] **M 100.1, m 68°, pK<sup>25</sup> -5.4 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Crystallise the acid from hot water or petroleum ether (b 60-80°). [Beilstein **2** IV 1555.]

**Dimethyl adipate** [627-93-0] **M 174.2, m 9-11°, b 109°/10mm, 121-123°/20mm, 235°/760mm, d<sub>4</sub><sup>20</sup> 1.0642, n<sub>D</sub><sup>20</sup> 1.4292.** Dissolve it in Et<sub>2</sub>O, wash with NaHCO<sub>3</sub>, H<sub>2</sub>O, dry over MgSO<sub>4</sub>, filter, evaporate and distil it several times until the IR and NMR are consistent with the structure [Lorette & Brown *J Org Chem* **24** 261 1959, Hoffmann & Weiss *J Am Chem Soc* **79** 4759 1957]. [Beilstein **2** IV 1959.]

**Dimethyl adipimidate dihydrochloride** [14620-72-5] **M 245.1, m 218-220°, 222-224°.** If the salt smells of HCl, then wash it with MeOH and dry Et<sub>2</sub>O (1:3) under N<sub>2</sub> until the free HCl is completely removed. Recrystallise it from MeOH/Et<sub>2</sub>O (it is very important that the solvents are super dry) [Hartman & Wold *Biochemistry* **6** 2439 1967, McElvain & Shroeder *J Am Chem Soc* **71** 40 1949].

**Dimethylamine** [124-40-3] **M 45.1, f -92.2°, b 0°/563mm, 6.9°/760mm, pK<sup>25</sup> 10.73.** Dry dimethylamine by passage through a KOH-filled tower, or by standing with sodium pellets at 0° during 18 hours. It is a strong base — *do not inhale its vapours*. [Beilstein **4** IV 128.]

§ A dimethylaminomethyl polystyrene supported version is commercially available.

**Dimethylamine hydrochloride** [506-59-2] **M 81.6, m 171°.** Crystallise the salt from hot CHCl<sub>3</sub> or absolute EtOH. It also recrystallises from MeOH/ether solution. Dry it in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>, then P<sub>2</sub>O<sub>5</sub>. *Hygroscopic*. [Beilstein **4** IV 132.]

**2-Dimethylaminoethanol** [108-01-0] **M 89.1, b 134.5-135.5°, d<sub>4</sub><sup>20</sup> 1.430, n<sub>D</sub><sup>20</sup> 1.4362, pK<sup>25</sup> 9.23.** Dry the amine with anhydrous K<sub>2</sub>CO<sub>3</sub> or KOH, and fractionally distil it. [Beilstein **4** IV 1424.]

**1-(3-Dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDCI, DEC, 1-ethyl-3-(3-dimethylaminopropyl) carbodiimide hydrochloride)** [25952-53-8] **M 191.7, m 113.5-114.5°, 114-116°, pK<sub>Est</sub> ~ 10.3.** It is an excellent H<sub>2</sub>O-soluble peptide coupling reagent. It is purified by dissolving (*ca* 1g) in CH<sub>2</sub>Cl<sub>2</sub> (10ml) at room temperature and then add dry Et<sub>2</sub>O (~110ml) dropwise and the crystals that separate are collected, washed with dry Et<sub>2</sub>O, recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O and dried in a vacuum over P<sub>2</sub>O<sub>5</sub>. It is important to work in a dry atmosphere or work rapidly, and then dry the solid as soon as possible. The material is moderately *hygroscopic*, but once it becomes wet it reacts slowly with H<sub>2</sub>O. Store it away from moisture at -20° to slow down the hydrolysis process. The *free base* has **b** 47-48°/0.27mm, 53-54°/0.6mm, n<sub>D</sub><sup>25</sup> 1.4582. The *methiodide* is recrystallised from CHCl<sub>3</sub>/EtOAc, the crystals are filtered off, washed with dry Et<sub>2</sub>O, recrystallised from CHCl<sub>3</sub>/Et<sub>2</sub>O, and dried *in vacuo* over P<sub>2</sub>O<sub>5</sub>, **m** 93-95°, 94-95°. [Sheehan et al. *J Am Chem Soc* **87** 2492 1965, Sheehan & Cruickshank *Org Synth Coll Vol V* 555 1973.]

§ A polymer bound version is commercially available.

***N,N*-Dimethylbiuret** [7710-35-2] **M 131.1, m 178°.** Purify it by repeated crystallisation from the melt, or from H<sub>2</sub>O. [Bredereck & Richter *Chem Ber* **99** 2461 1968, Dunning & Close *J Am Chem Soc* **75** 3615 1953.]

**2,3-Dimethyl-1,3-butadiene** [513-81-5] **M 82.2, m -69-70°, b 68-69°/760mm, d<sub>4</sub><sup>20</sup> 0.727, n<sub>D</sub><sup>20</sup> 1.4385.** Distil

it from  $\text{NaBH}_4$ , and purify it by zone melting. [*Beilstein* 1 IV 1023.]

**1,3-Dimethylbutadiene sulfone (1,3-dimethylsulfolene, 2,4-dimethyl-2,5-dihydrothiophen-1,1-dioxide)** [10033-92-8] **M 145.2, m 40.4-41.0°**. Crystallise the sulfone from diethyl ether (three times), or from  $\text{CCl}_4$  (**m 39.5-40°**). [Grummitt et al. *J Am Chem Soc* 72 5768 1950, *Beilstein* 17/1 V 97.]

**2,2-Dimethylbutane** [75-83-2] **M 86.2, b 49.7°, d<sub>4</sub><sup>20</sup> 0.649, n<sub>D</sub><sup>25</sup> 1.36595**. Distil it azeotropically with MeOH, then wash it with water, dry ( $\text{Na}_2\text{SO}_4$ ) it, and distil it. [*Beilstein* 1 IV 367.]

**2,3-Dimethylbutane** [79-29-8] **M 86.2, b 58.0°, d<sub>4</sub><sup>20</sup> 1.375, n<sub>D</sub><sup>25</sup> 1.37231**. Distil it from sodium, pass it through a column of silica gel (activated by heating in nitrogen to 350° before use) to remove unsaturated impurities, and again distil it from sodium. Also distil it azeotropically with MeOH, then wash with water, dry ( $\text{Na}_2\text{SO}_4$ ), filter and redistil it. [*Beilstein* 1 IV 371.]

**2,3-Dimethylbut-2-ene** [563-79-1] **M 84.2, b 72-73°/760mm, d<sub>4</sub><sup>20</sup> 0.708, n<sub>D</sub><sup>20</sup> 1.41153**. Purify it by GLC on a column of 20% squalene on chromosorb P at 50° [Flowers & Rabinovitch *J Phys Chem* 89 563 1985]. Also wash it with 1M NaOH solution followed by  $\text{H}_2\text{O}$ . Dry it over  $\text{Na}_2\text{SO}_4$ , distil it over powdered KOH under nitrogen and pass it through activated alumina before use. [Woon et al. *J Am Chem Soc* 108 7990 1986, Wong et al. *J Am Chem Soc* 109 3428 1987, *Beilstein* 1 IV 853.]

**Dimethylcarbamoyl chloride** [79-44-7] **M 107.5, m -33°, b 34°/0.1mm, d<sub>4</sub><sup>20</sup> 1.172, n<sub>D</sub><sup>20</sup> 1.4511**. It must be distilled under high vacuum to avoid decomposition. It is moisture sensitive. [*Beilstein* 4 IV 224.]

**Dimethyl carbonate** [616-38-6] **M 90.1, m 2-4°, 4.65°, b 89.5°/755mm, 90.2°/atm, 90-91°/atm, d<sub>4</sub><sup>20</sup> 1.079, n<sub>D</sub><sup>20</sup> 1.3687**. If the reagent has broad intense bands at  $3300\text{cm}^{-1}$  and above (i.e. OH stretching), then it should be purified further. It will contain small amounts of water and/or alcohol which form azeotropes with it. Wash it successively with 10%  $\text{Na}_2\text{CO}_3$  solution, saturated  $\text{CaCl}_2$ ,  $\text{H}_2\text{O}$ , and dry it by shaking mechanically for 1 hour with anhydrous  $\text{CaCl}_2$ , and fractionate. Alternatively, stand it for several days in contact with Linde type 4A molecular sieves, then fractionally distil it. The middle fraction is frozen slowly at 2°, several times, retaining 80% of the liquid at each cycle. [Bowden & Butler *J Chem Soc* 78 1939, Vogel *J Chem Soc* 1847 1948, *Beilstein* 3 IV 3.]

**Dimethyl dicarbonate (dimethyl pyrocarbonate)** [4525-33-1] **M 134.1, m 15.2°, b 45-46°/5mm, d<sub>4</sub><sup>20</sup> 1.2585, n<sub>D</sub><sup>20</sup> 1.3950**. If the reagent has broad intense bands at  $3300\text{cm}^{-1}$  and above (i.e. OH stretching), then it should be purified further. Dissolve it in  $\text{Et}_2\text{O}$ , shake this with a small volume of 0.1N HCl, dry the  $\text{Et}_2\text{O}$  solution ( $\text{Na}_2\text{SO}_4$ ), and distil it *in vacuo* at as low pressure as possible, but below 100°, to give a clear liquid. It decomposes to  $\text{CO}_2$  and dimethyl carbonate on heating at 123-149°. It is readily hydrolysed by  $\text{H}_2\text{O}$  and is a yeast inhibitor. It is an **IRRITANT**. [Brysov et al. *J Org Chem USSR* 10 2551 1974, Boehm & Mehta *Chem Ber* 71 1797 1938, *Beilstein* 3 IV 17.]

**1,5-Dimethyl-1,5-diazaundecamethylene polymethobromide (Hexadimethrene, Polybrene)** [28728-55-4] **M 5000—10,000 polymer**. Purify it by chromatography on Dowex 50 and/or by filtration through alumina before use [Frank Hoppe-Seyler's *Z Physiol Chemie* 360 997 1979]. It is hygroscopic, and its solubility in  $\text{H}_2\text{O}$  is 10%.

***N,N'*-Dimethylethylenediamine [1,2-bis(methylamino)ethane]** [110-70-3] **M 88.2, b 110-112°/750mm, 119°/760mm, d<sub>4</sub><sup>20</sup> 0.819, n<sub>D</sub><sup>20</sup> 1.431, (pK<sup>25</sup> 7.01 and 9.88)**. This strong base has been prepared in various ways including hydrolysis of *N,N'*-di(benzylsulfonyl)-*N,N'*-dimethylethylenediamine (**m 217-219°**, from AcOH) and concentrated HCl (120-130°) [Johnson & Bailey *J Am Chem Soc* 38 2135 1916], the reaction of 1,2-dibromoethane with 21-33% aqueous  $\text{MeNH}_2$  in EtOH (reflux for 2 hours, 50% yield) [Kermack & Wight *J Chem Soc* 1421 1935, Woodburn & O'Gee *J Org Chem* 17 1235 1952], hydrolysis of *N,N'*-di(*p*-toluenesulfonyl)

*N,N'*-dimethylethylenediamine (1mol, **m 164°**, from AcOH) with H<sub>2</sub>SO<sub>4</sub> (8.2mol) and H<sub>2</sub>O (9mol) at 140-145°/7 hours [Boon *J Chem Soc* 307 1947], LAH reduction in THF of 3-benzenesulfonyloxy-5,6-dihydrouacil (**m 175-176°**, from iso-PrOH) and isolated as the *dibenzoyl derivative* (41%, **m 177-178°**, from \*C<sub>6</sub>H<sub>6</sub>) [Bauer *J Am Chem Soc* 78 1945 1956], and by catalytic hydrogenation of *N-benzyl-N,N'*-dimethylethylenediamine (**b 73-74°/0.1mm**) with 10% Pd/C [Jucker & Rissi *Helv Chim Acta* 45 2383 1962]. General isolation and purification procedures involve steam distillation of the product diamine, acidifying the distillate with HCl, evaporating to dryness, treating the residual salt with 10% excess of cold 32% aqueous NaOH, collecting the organic layer (**care**: very caustic solution), drying it over solid KOH, and distilling from it. Final distillation of the dry base over Na has been reported (see also [107-15-3]). The diamine is a strong base and readily absorbs CO<sub>2</sub> and H<sub>2</sub>O from the atmosphere—store it in a dark stoppered bottle, wax the stopper if it should be stored for long periods. The *dihydrochloride* [5752-40-9] has **m 235-236° (dec)**, and the *picrate* has **m 160°** (from EtOH and Me<sub>2</sub>CO in rectangular plates). It forms complexes with Cu, Ni, and Pt among other metals. Its FT-IR (neat) has  $\nu_{\max}$  at 3285.5, 2788.2, 1447.2, 1346.5, 1251.0, 1106.5, 1040.5, 876.9, 763.7 cm<sup>-1</sup>. [For pK and metal complexes see Gastafson & Martell *J Am Chem Soc* 81 525 1959, *Beilstein* 4 H 250, 4 I 415, 4 II 689, 4 III 512, 4 IV 1171.]

**Dimethyl disulfide** [624-92-0] **M 94.2, f -98°, b 40°/12mm, 110°/760mm, d<sub>4</sub><sup>20</sup> 1.0605, n 1.5260**. Pass it through neutral alumina before use. [Trost *Chem Rev* 78 363 1978, *Beilstein* 1 IV 1281.]

***N,N*-Dimethylformamide (DMF)** [68-12-2] **M 73.1, b 40°/10mm, 61°/30mm, 88°/100mm, 153°/760mm, d<sub>4</sub><sup>20</sup> 0.948, n<sub>D</sub><sup>25</sup> 1.4269, pK<sub>D</sub><sup>25</sup> -0.3**. DMF decomposes slightly at its normal boiling point to give small amounts of dimethylamine and carbon monoxide. The decomposition is catalysed by acidic or basic materials, so that even at room temperature DMF is appreciably decomposed if allowed to stand for several hours with solid KOH, NaOH or CaH<sub>2</sub>. If these reagents are used as dehydrating agents, therefore, they should not be refluxed with the DMF. Use of CaSO<sub>4</sub>, MgSO<sub>4</sub>, silica gel or Linde type 4A molecular sieves is preferable, followed by distillation under reduced pressure. This procedure is adequate for most laboratory purposes. Larger amounts of water can be removed by azeotropic distillation with \*benzene (10% v/v, previously dried over CaH<sub>2</sub>), at atmospheric pressure: water and \*benzene distil below 80°. The liquid remaining in the distillation flask is further dried by adding MgSO<sub>4</sub> (previously ignited overnight at 300-400°) to give 25g/L. After shaking for one day, a further quantity of MgSO<sub>4</sub> is added, and the DMF is distilled at 15-20mm pressure through a 3-ft vacuum-jacketed column packed with steel helices. However, MgSO<sub>4</sub> is an inefficient drying agent, leaving about 0.01M water in the final DMF. More efficient drying (to around 0.001-0.007M water) is achieved by standing with powdered BaO, followed by decanting before distillation, then with alumina powder (50g/L, previously heated overnight to 500-600°), and distilling from more of the alumina, or by refluxing at 120-140° for 24 hours with triphenylchlorosilane (5-10g/L), then distilling at *ca* 5mm pressure [Thomas & Rochow *J Am Chem Soc* 79 1843 1957]. Free amine in DMF can be detected by the colour reaction with 1-fluoro-2,4-dinitrobenzene. It has also been purified by drying overnight over KOH pellets and then distilling from BaO through a 10 cm Vigreux column [Jasiewicz et al. *Exp Cell Res* 100 213 1976]. [For efficiency of desiccants in drying dimethylformamide see Burfield & Smithers *J Org Chem* 43 3966 1978, and for a review on purification, tests of purity and physical properties, see Juillard *Pure Appl Chem* 49 885 1977.]

It has been purified by distilling from K<sub>2</sub>CO<sub>3</sub> under high vacuum and fractionated in an all-glass apparatus. The middle fraction is collected, degassed (seven or eight freeze-thaw cycles) and redistilled under as high a vacuum as possible [Mohammad & Kosower *J Am Chem Soc* 93 2713 1971]. [*Beilstein* 4 IV 171.]

**Rapid purification:** Stir over CaH<sub>2</sub> (5% w/v) overnight, filter, then distil at 20 mmHg. Store the distilled DMF over 3A or 4A molecular sieves. For solid phase synthesis, the DMF used must be of high quality and free from amines.

***d,l*-2,4-Dimethylglutaric acid** [2121-67-7] **M 160.2, m 144-145° pK<sub>Est(1)</sub> ~4.4, pK<sub>Est(2)</sub> ~5.4**. Distil the acid in steam and recrystallise it from ether/petroleum ether. [*Beilstein* 2 H 681, 2 I 284, 2 II 590, 21 III 1755, 2 IV 2022.]

**3,3-Dimethylglutaric acid** [4839-46-7] **M 160.2, m 103-104°, b 89-90°/2mm, 126-127°/4.5mm, pK<sub>1</sub><sup>25</sup> 3.85, pK<sub>2</sub><sup>25</sup> 6.45**. Crystallise the acid from water, \*benzene or ether/petroleum ether. Dry it in a vacuum. [*Beilstein*

21 H 391, 21 I 334, 21 II 309, 21 III/IV 4601, 21/9 V 592.]

**Dimethylglyoxime** [95-45-4] M 116.1, m 240°, pK<sub>1</sub><sup>25</sup> 10.60, pK<sub>2</sub><sup>25</sup> 11.85. Crystallise it from EtOH (10ml/g) or aqueous EtOH. [Beilstein 1 III 3105.] **TOXIC.**

**2,5-Dimethyl-2,4-hexadiene** [764-13-6] M 110.2, f 14.5°, b 132-134°, d<sub>4</sub><sup>20</sup> 0.773, n<sub>D</sub><sup>20</sup> 1.4796. Distil it, then repeatedly fractionally crystallise it by partial freezing. Immediately before use, the diene is passed through a column containing Woelm silica gel (activity I) and Woelm alumina (neutral) in separate layers. [Beilstein 1 IV 1043.]

**2,2-Dimethylhexane** [590-73-8] M 114.2, m -121.2°, b 107°, d<sub>4</sub><sup>20</sup> 0.695. Dry the hexane over type 4A molecular sieves and distil it. [Beilstein 1 IV 432.]

**2,5-Dimethylhexane** [592-13-2] M 114.2, m -91.2°, b 109°, d<sub>4</sub><sup>20</sup> 0.694. Dry the hexane over type 4A molecular sieves and distil it. [Beilstein 1 IV 434.]

**2,5-Dimethylhexane-2,5-diol** [110-03-2] M 146.2, m 88-90°. Purify the diol by fractional crystallisation. Then the diol is dissolved in hot acetone, treated with activated charcoal, and filtered while hot. The solution is cooled and the diol is filtered off and washed well with cold acetone. The crystallisation process is repeated several times, and the crystals are dried under a vacuum in a freeze-drying apparatus [Goates et al. *J Chem Soc, Faraday Trans 1* 78 3045 1982]. [Beilstein 1 IV 2600.]

**1,1-Dimethylhydrazine** [57-14-7] M 60.1, b 60.1°/702mm, d<sub>4</sub><sup>20</sup> 0.790, n<sub>D</sub><sup>20</sup> 1.408 pK<sup>30</sup> 7.21. Fractionally distil the hydrazine through a 4-ft column packed with glass helices. Precipitate it as its oxalate from diethyl ether solution. After crystallising from 95% EtOH, the salt is decomposed with aqueous saturated NaOH, and the free base is distilled, dried over BaO and redistilled [McBride & Kruse *J Am Chem Soc* 79 572 1957]. Distillation and storage should be under nitrogen. [Beilstein 4 IV 3322.]

**Dimethyl itaconate** [617-52-7] M 158.2, m 38°, b 208°, d 1.124. Crystallise the ester from MeOH by cooling to -78°. [Beilstein 2 IV 2229.]

**Dimethylmaleic anhydride** [766-39-2] M 126.1, m 96°, b 225°/760mm. Distil the anhydride from \*benzene/ligroin and sublime in a vacuum. [Beilstein 17/11 V 69.]

**Dimethylmalonic acid** [595-46-0] M 132.1, m 192-193° pK<sub>1</sub><sup>25</sup> 3.03, pK<sub>2</sub><sup>25</sup> 5.73. Crystallise the acid from \*benzene/petroleum ether. It sublimes in a vacuum with slight decomposition. [Beilstein 2 IV 1955.]

**Dimethylnitrosamine (N-nitrosodimethylamine)** [62-75-9] M 74.0, m -28°, b 149-150°/atm, 153°/774mm, d<sub>4</sub><sup>20</sup> 1.006, n<sub>D</sub><sup>20</sup> 1.4370. Dry the nitrosamine over anhydrous K<sub>2</sub>CO<sub>3</sub> or dissolve it in Et<sub>2</sub>O, dry it over solid KOH, filter, evaporate Et<sub>2</sub>O and distil the yellow oily residue through a 30cm fractionating column discarding the first fraction which may contain Me<sub>2</sub>N. Also dry over CaCl<sub>2</sub> and distil it at atmospheric pressure. All operations should be done in an efficient fume cupboard as the vapors are **TOXIC** and **CARCINOGENIC**. [Fischer *Chem Ber* 8 1588 1875, Romberg *Rec Trav Chim Pays Bas* 5 248 1886, Hatt *Org Synth Coll Vol II* 211 1961, Krebs & Mandt *Chem Ber* 108 1130 1975.]

**2,6-Dimethyl-2,4,6-octatriene** see *neo*-alloocimene below.

**Dimethylolurea (di-[hydroxymethyl]urea)** [140-95-4] M 120.1, m 137-139°. Crystallise it from aqueous 75% EtOH. [Beilstein 3 IV 107.]

**Dimethyl oxalate** [553-90-2] M 118.1, m 54°, b 163-165°, d<sub>4</sub><sup>20</sup> 1.148. Crystallise the ester repeatedly from EtOH. De-gas it under nitrogen at high vacuum and distil it. [Beilstein 2 IV 1847.]

**2,4-Dimethylpentane** [108-08-7] M 100.2, b 80.5°,  $d_4^{20}$  0.763,  $n_D^{20}$  1.3814,  $n_D^{25}$  1.37882. Extract it repeatedly with conc H<sub>2</sub>SO<sub>4</sub>, wash with water, dry and distil it. *Alternatively*, percolated it through silica gel (previously heated in nitrogen to 350°). Purify it by azeotropic distillation with EtOH, followed by washing out the EtOH with water, drying and distilling. [Beilstein 1 IV 406.]

**4,4-Dimethyl-1-pentene** [762-62-9] M 98.2, b 72.5°/760mm,  $d_4^{20}$  0.6827,  $n_D^{20}$  1.3918. Purify it by passing through alumina before use [Traylor et al. *J Am Chem Soc* 109 3625 1987]. [Beilstein 1 IV 869.]

**Dimethyl peroxide** [690-02-8] M 62.1, b 13.5°/760mm,  $d_4^{20}$  0.8677,  $n_D^{20}$  1.3503. Purify dimethyl peroxide by repeated trap-to-trap fractionation until no impurities could be detected by gas IR spectroscopy [Haas & Oberhammer *J Am Chem Soc* 106 6146 1984]. *All necessary precautions should be taken in case of EXPLOSION.*

**2,2-Dimethyl-1,3-propanediol (neopentyl glycol)** [126-30-7] M 104.2, m 128.4-129.4°, b 208°/760mm. Crystallise the diol from \*benzene or acetone/water (1:1). [Beilstein 1 IV 2551.]

**2,2-Dimethyl-1-propanol (neo-pentyl alcohol)** [75-84-3] M 88.2, m 52°, b 113.1°/760mm. It is difficult to distil because it is a solid at ambient temperatures. Purify it by fractional crystallisation and sublimation. [Beilstein 1 IV 1690.]

***N,N*-Dimethylpropionamide** [758-96-3] M 101.2, b 175-178°,  $d_4^{20}$  0.920,  $n_D^{20}$  1.440. Shake the amide over BaO for 1-2 days, then distil it under reduced pressure. [Beilstein 4 III 126.]

***meso*- $\alpha,\beta$ -Dimethylsuccinic acid** [608-40-2] M 146.1, m 211°,  $pK_1^{25}$  3.77,  $pK_2^{25}$  5.36. Crystallise the *meso*-acid from EtOH/ether or EtOH/chloroform.

**2,2-Dimethylsuccinic acid** [597-43-3] M 146.1, m 141°,  $pK_1^{20}$  4.15,  $pK_2^{20}$  6.40. Crystallise the acid from EtOH/ether or EtOH/chloroform. [Beilstein 2 IV 1996.]

**( $\pm$ )-2,3-Dimethylsuccinic acid** [13545-04-5] M 146.1, m 129°,  $pK_1^{25}$  3.82,  $pK_2^{25}$  5.98. Crystallise the *racemic* acid from water. [Beilstein 2 IV 1998.]

**Dimethyl sulfide** [75-18-3] M 62.1, f -98.27°, b 0°/172mm, 37.5-38°/760mm,  $d_4^{21}$  0.8458,  $n_D^{25}$  1.4319. Purify dimethyl sulfide *via* the Hg(II) chloride complex by dissolving 1 mole of Hg(II)Cl<sub>2</sub> in 1250ml of EtOH and slowly adding the boiling alcoholic solution of Me<sub>2</sub>S to give the right ratio for 2(CH<sub>3</sub>)<sub>2</sub>S.3HgCl<sub>2</sub>. After recrystallisation of the complex to constant melting point, 500g of complex is heated with 250ml conc HCl in 750ml of water. The sulfide is separated, washed with water, and dried with CaCl<sub>2</sub> and CaSO<sub>4</sub>. Finally, it is distilled under reduced pressure from sodium. Precautions should be taken (*efficient fume hood*) because of its very UNPLEASANT ODOUR and TOXICITY. [Beilstein 1 IV 1275.]

**2,4-Dimethylsulfolane** [1003-78-7] M 148.2, b 128°/77mm,  $d_4^{25}$  1.1314,  $n_D^{20}$  1.474. Distil the sulfolane in a vacuum. [Beilstein 17/1 IV 97.]

**Dimethyl sulfone** [67-71-0] M 94.1, m 109°. Crystallise the sulfone from water. Dry it over P<sub>2</sub>O<sub>5</sub>. [Beilstein 1 IV 1279.]

**Dimethyl sulfoxide (DMSO)** [67-68-5] M 78.1, m 18.0-18.5°, b 75.6-75.8°/12mm, 190°/760mm,  $d_4^{20}$  1.100,  $n_D^{20}$  1.479. This colourless, odourless, very *hygroscopic* liquid, is synthesised from dimethyl sulfide. The main impurity is water, with a trace of dimethyl sulfone. The Karl-Fischer test is applicable. It is dried with Linde types 4A or 13X molecular sieves, by prolonged contact and passage through a column of the material, then distilled under reduced pressure. Other drying agents include CaH<sub>2</sub>, CaO, BaO and CaSO<sub>4</sub>. It can also be fractionally crystallised by partial freezing. More extensive purification is achieved by standing overnight with freshly heated and cooled chromatographic grade alumina. It is then refluxed for 4 hours over CaO, dried

over  $\text{CaH}_2$ , and fractionally distilled at low pressure. For efficiency of desiccants in drying dimethyl sulfoxide see Burfield and Smithers [*J Org Chem* **43** 3966 1978, Sato et al. *J Chem Soc, Dalton Trans* 1949 1986]. [Reddy *Pure Appl Chem* **25** 459 1969, *Beilstein* **1** IV 1277.]

**Rapid purification:** Stand over freshly activated alumina, BaO or  $\text{CaSO}_4$  overnight. Filter and distil it over  $\text{CaH}_2$  under reduced pressure (~12 mmHg). Store it over 4A molecular sieves.

***N,N*-Dimethylthiocarbamoyl chloride** [16420-13-6] **M 123.6, m 42-43°, b 64-65°/0.1mm.** Crystallise it twice from pentane and/or distil it at low pressure. [*Beilstein* **4** III 147.]

***sym*-Dimethylurea** [96-31-1] **M 88.1, m 106°.** Crystallise the urea from acetone/diethyl ether by cooling in an ice bath. Also crystallise it from EtOH and dry it at 50°/5mm for 24 hours [Bloemendahl & Somsen *J Am Chem Soc* **107** 3426 1985]. [*Beilstein* **4** IV 207.]

**2,2-Dinitropropane** [595-49-3] **M 162.1, m 53.5°.** Crystallise it from EtOH or MeOH. Dry it over  $\text{CaCl}_2$  or *in vacuo* for 1 hour just above the melting point. [*Beilstein* **2** H 117, **2** II 79, **2** III 261, **2** IV 234.]

**Diocetadecyldimethylammonium bromide** [3700-67-2] **M 630.9, m 161-163°.** Crystallise the bromide from acetone, then MeOH [Lukac *J Am Chem Soc* **106** 4387 1984]. Also purify it by chromatography on alumina by washing with  $^*\text{C}_6\text{H}_6$  and eluting with  $\text{Me}_2\text{CO}$ , evaporating and crystallising from MeCN [Swain & Kreevoy *J Am Chem Soc* **77** 1126 1955]. [*Beilstein* **4** IV 829.]

***N,N*-Diocetadecyl methylamine (hydrogen ionophore III)** [4088-22-6] **M 536.0, m 40°, 44-46°, 48-49°, b 252-259°,  $\text{pK}_{\text{Est}} \sim 10$ .** It can be distilled at high vacuum, but dissolving in  $^*\text{C}_6\text{H}_6$ , filtering and evaporating give a waxy solid suitable for electrode use. It recrystallises from  $\text{Me}_2\text{CO}$  or MeCN. [Hoerr et al. *J Org Chem* **9** 201 1944, Wu & Yu *Talanta* **34** 577 1987, *Beilstein* **4** III 435.]

**Dipropylene glycol (octan-4,5-diol)** [110-98-5] **M 134.2, b 109-110°/8mm,  $d_4^{20}$  1.022,  $n_D^{20}$  1.441.** Fractionally distil the diol below 15mm pressure, using a packed column and taking precautions to avoid absorption of water. [*Beilstein* **1** IV 2473.]

**Di-*n*-propyl ketone (4-pentanone)** [123-19-3] **M 114.2, b 143.5°,  $d_4^{20}$  0.8143,  $n_D^{20}$  1.40732.** Dry 4-pentanone with  $\text{CaSO}_4$ , then distil it from  $\text{P}_2\text{O}_5$  under nitrogen. [*Beilstein* **1** IV 3323.]

**Di-*n*-propyl sulfide** [111-47-7] **M 118.2, b 141-142°,  $d_4^{20}$  0.870,  $n_D^{20}$  1.449.** Wash the sulfide with aqueous 5% NaOH, then water. Dry it with  $\text{CaCl}_2$  and distil it from Na [Dunstan & Griffiths *J Chem Soc* 1344 1962]. [*Beilstein* **1** IV 1452.]

***S*-1,2-Distearin (*S*-glycerol-1,2-distearate)** [*S*- 1429-59-0, *RS*- 51063-97-9, *R*- 1188-58-5] **M 625.0, m 76-77°,  $[\alpha]_D^{20} -2.8^\circ$  (c 6.3,  $\text{CHCl}_3$ ),  $[\alpha]_{546}^{20} +1.4^\circ$  (c 10,  $\text{CHCl}_3/\text{MeOH}$ , 9:1).** Crystallise the glyceride from chloroform/petroleum ether. [cf p 678, *Beilstein* **2** IV 1231.]

**1,4-Dithioerythritol (DTE, erythro-2,3-dihydroxy-1,4-dithiobutane)** [6892-68-8] **M 154.3, m 82-84°,  $\text{pK}_1$  9.0,  $\text{pK}_2$  9.9.** Crystallise DTE from ether/hexane and store it in the dark at 0°. [*Beilstein* **1** III 2360.]

**Dithiooxamide (rubeanic acid)** [79-40-3] **M 120.2, m >300°.** Crystallise dithiooxamide from EtOH and sublime it at high vacuum. [*Beilstein* **2** IV 1871.]

***RS*-1,4-Dithiothreitol (DTT, Cleland's reagent)** [27565-41-9] **M 154.3, m 42-43°,  $\text{pK}_1$  8.3,  $\text{pK}_2$  9.5.** Crystallise DTT from ether and sublime it at 37°/0.005mm. It should be stored at 0°. [*Beilstein* **1** III 2360.]

***All-cis*-4,7,10,13,16,19-Docosaehaenoic acid** [6217-54-5] **M 328.5, m -44.1°,  $n_D^{20}$  1.5017,  $\text{pK}_{\text{Est}} \sim 4.6$ .** Its solubility in  $\text{CHCl}_3$  is 5%. It has been purified from fish oil by GLC using Ar as mobile phase and EGA as stationary phase with an ionisation detector [UV: Stoffel & Ahrens *J Lipid Res* **1** 139 1959], and *via* the ester by evaporative "molecular" distillation using a 'continuous molecular still' at  $10^{-4}$  mm with the highest temperature

being 110° and a total contact time with the hot surface being 60 seconds [Farmer & van den Heuvel *J Chem Soc* 427 1938]. The *methyl ester* [2566-90-7] has **b** 208-211°/2mm,  $d_4^{20}$  0.9398,  $n_D^{20}$  1.5035. With Br<sub>2</sub> it forms a *dodecaboromide* **m ca** 240°(dec). Also, the acid was converted to the methyl ester and purified through a three-stage molecular still [as described by Sutton *Chem Ind (London)* 11383 1953] at 96°, and the rate was adjusted so that one-third of the material was removed each cycle of three distillations. The distillate (numbered 4) (13g) was dissolved in EtOH (100ml containing 8g of KOH) at -70° and set aside for 4 hours at 30° with occasional shaking under a vacuum. Water (100ml) was added and the solution was extracted with pentane, washed with HCl, dried (MgSO<sub>4</sub>), filtered and evaporated to give a clear oil (11.5g) **m** -44.5° to -44.1°. In the catalytic hydrogenation of the oil six mols of H<sub>2</sub> are absorbed and *docosanoic acid* (*behenic acid*) is produced with **m** 79.0-79.3° undepressed with an authentic sample (see docosanoic acid below) [Whitcutt *Biochem J* 67 60 1957]. [Beilstein 2 IV 1812.]

**Docosane (C22)** [629-97-0] **M 310.6, m 47°, b 224°/15mm.** Crystallise docosane from EtOH or ether. [Beilstein 1 IV 572.]

**Docosanoic acid (behenic acid)** [112-85-6] **M 340.6, m 81-82°, pK<sub>Est</sub> ~4.9.** Crystallise the acid from ligroin. [Francis & Piper *J Am Chem Soc* 61 577 1939, Beilstein 2 IV 1290.]

**1-Docosanol (behenyl alcohol)** [661-19-8] **M 182.3, m 70.8°.** Crystallise docosanol from ether or chloroform/ether. [Beilstein 1 IV 1906.]

***n*-Dodecane** [112-40-3] **M 170.3, b 97.5-99.5°/5mm, 216°/760mm,  $d_4^{20}$  0.748,  $n_D^{20}$  1.42156.** Pass it through a column of Linde type 13X molecular sieves. Store it in contact with, and distil it from sodium. Pass it through a column of activated silica gel. It has been crystallised from diethyl ether at -60°. Unsaturated dry material which remained after passage through silica gel has been removed by catalytic hydrogenation (Pt<sub>2</sub>O) at 45lb/in<sup>2</sup> (3.06 atmospheres), followed by fractional distillation under reduced pressure [Zook & Goldey *J Am Chem Soc* 75 3975 1953]. It has also purified by partial crystallisation from the melt. [Beilstein 1 IV 498.]

**Dodecane-1,10-dioic acid (decane-1,10-dicarboxylic acid)** [693-23-2] **M 230.3, m 129°, b 245°/10mm, pK<sub>Est</sub> ~4.8.** Crystallise the dioic acid from water, 75% or 95% EtOH (solubility is 10%), or glacial acetic acid. [Beilstein 2 IV 2126.]

**1-Dodecanol (dodecyl alcohol)** [112-53-8] **M 186.3, m 24°, b 91°/1mm, 135°/10mm, 167°/40mm, 213°/200mm, 259°/760mm,  $d_4^{24}$  0.8309 (liquid).** Crystallise dodecanol from aqueous EtOH, and distil it through a spinning-band column under vacuum. [Ford & Marvel *Org Synth* 10 62 1930, Beilstein 1 IV 1844.]

**1-Dodecanthiol** [112-55-0] **M 202.4, b 111-112°/3mm, 153-155°/24mm,  $d_4^{20}$  0.844,  $n_D^{20}$  1.458, pK<sub>Est</sub> ~10.8.** Dry it with CaO for several days, then distil it from CaO. [Beilstein 1 IV 1851.]

**Dodecylamine** [124-22-1] **M 185.4, m 28°, 27-29°, 120-121°/2mm, 134°/15mm, 156°/33mm, pK<sup>25</sup> 10.63.** Fractionally distil the amine, preferably under N<sub>2</sub> and in a vacuum. Store it in the absence of CO<sub>2</sub>. It can be recrystallised from *n*-hexane at low temperature. The *hydrochloride* crystallises from Me<sub>2</sub>CO (**m** 182-183°) or CHCl<sub>3</sub>/petroleum ether (**m** 185-187°). [Magnien & Baltzly *J Org Chem* 23 2029 1958, Beilstein 4 H 200, III 406, 4 IV 794.]

**Dodecylammonium butyrate** [17615-97-3] **M 273.4, m 39-40°, 39-41°, pK<sup>25</sup> 10.63 (for free base).** Recrystallise the salt from *n*-hexane. [Beilstein 4 III 409, 4 IV 791.]

**Dodecylammonium propionate** [17448-65-6] **M 259.4, m 55-56°.** Recrystallise the salt from hexanol/petroleum ether (b 60-80°). [Beilstein 4 III 409, 4 IV 797.]

**Dodecyldimethylamine oxide** [1643-20-5] **M 229.4, m 102°.** Crystallise the oxide from acetone or ethyl

acetate. [Bunton et al. *J Org Chem* **52** 3832 1987, *Beilstein* **4** III 410, **4** IV 798.]

**Dodecyl ether (didodecyl ether)** [4542-57-8] **M 354.6, m 32.5-33°, 33°, b 175°/0.15mm, d<sup>36</sup> 0.8127, n<sup>39</sup> 1.4393.** Distil the ether in a vacuum, then crystallise it from MeOH or MeOH/\*benzene. [Mannich & Nadelmann *Chem Ber* **63** 799 1930. Butterworth & Hey *J Chem Soc* 390 1940, *Beilstein* **1** III 1785, **1** IV 1846.]

**Dodecyl methacrylate (lauryl methacrylate)** [142-90-5] **M 254.4, m -7°, b 142°/4mm, d<sup>25</sup> 0.8717, n<sup>25</sup> 1.4330.** Purify the ester by fractional distillation in a high vacuum. Add 0.05% of hydroquinone monomethyl ether as stabiliser. [Rehberg & Fischer *Ind Eng Chem* **40** 1430 1948, *Beilstein* **2** III 1290, **2** IV 1528.]

**Dodecyltrimethylammonium bromide** [1119-94-4] **M 308.4, m 246°(dec).** Purify the salt by repeated crystallisation from acetone. Wash it with diethyl ether and dry it in a vacuum oven at 60° [Dearden & Wooley *J Phys Chem* **91** 2404 1987]. [*Beilstein* **4** IV 798.]

**Dodecyltrimethylammonium chloride** [112-00-5] **M 263.9, m 246°(dec).** Dissolve the chloride in MeOH, treat with active charcoal, filter and dry it *in vacuo* [Waldenburg *J Phys Chem* **88** 1655 1984], or recrystallise it several times from 10% EtOH in acetone. It has also been repeatedly crystallised from EtOH/ether or MeOH. [Cella et al. *J Am Chem Soc* **74** 2062 1952, *Beilstein* **4** IV 79.]

**Eicosane (C<sub>20</sub>)** [112-95-8] **M 282.6, m 36-37°, b 205°/15mm, d<sup>36.7</sup> 0.7779, n<sup>40</sup> 1.43453.** Crystallise eicosane from EtOH. [*Beilstein* **1** IV 563.]

**Elaidic (trans-oleic) acid** [112-79-8] **M 282.5, m 44.5°, pK<sup>25</sup> 4.9.** Crystallise the acid from acetic acid, then EtOH. [*Beilstein* **2** IV 1647.]

**RS-Epichlorohydrin (± 2-chloromethyloxirane)** [106-89-8] **M 92.5, b 115.5°, d<sub>4</sub><sup>20</sup> 1.180, n<sub>D</sub><sup>20</sup> 1.438.** Distil epichlorohydrin under atmospheric pressure, heat it on a steam bath with one-quarter its weight of CaO, then decant and fractionally distil it. [*Beilstein* **17** V 20.]

**1,2-Epoxybutane** [106-88-7] **M 72.1, b 66.4-66.6°, d<sub>4</sub><sup>20</sup> 0.837, n<sub>D</sub><sup>20</sup> 1.3841.** Dry it with CaSO<sub>4</sub>, and fractionally distil it through a long (126cm) glass helices-packed column. The first fraction contains a water azeotrope. [*Beilstein* **17** II 17.]

**Erucic acid (cis-13-docosenoic acid)** [112-86-7] **M 338.6, m 33.8°, b 358°/400mm, pK<sub>Est</sub> ~4.9.** Crystallise erucic acid from MeOH. [*Beilstein* **2** IV 1676.]

**Ethane** [74-84-0] **M 30.1, f -172°, b -88°, d<sub>4</sub><sup>0</sup> 1.0493 (air = 1).** Ethylene can be removed by passing the gas through a sintered-glass disc into fuming H<sub>2</sub>SO<sub>4</sub> then slowly through a column of charcoal saturated with bromine. Bromine and HBr are removed by passage through firebrick coated with *N,N*-dimethyl-*p*-toluidine. The ethane is also passed over KOH pellets (to remove CO<sub>2</sub>) and dried with Mg(ClO<sub>4</sub>)<sub>2</sub>. Further purification is by several distillations of liquified ethane, using a condensing temperature of -195°. Yang and Gant [*J Phys Chem* **65** 1861 1961] treated ethane by standing it for 24 hours at room temperature in a steel bomb with activated charcoal treated with bromine. They then immersed the bomb in a Dry-ice/acetone bath and transferred the ethane to an activated charcoal trap cooled in liquid nitrogen. (The charcoal had previously been degassed by pumping for 24 hours at 450°.) By allowing the trap to warm slowly, the ethane distils, and only the middle third fraction is kept. Removal of methane is achieved using Linde type 13X molecular sieves (previously degassed by pumping for 24 hours at 450°) in a trap which, after cooling in Dry-ice/acetone, is saturated with ethane. After pumping for 10 minutes, the ethane is recovered by warming the trap to 25°. (The final gas contains less than 10<sup>-4</sup> mole % of either ethylene or methane). [*Beilstein* **1** IV 108.]

**Ethanesulfonyl chloride** [594-44-5] **M 128.6, b 55°/9mm, 62°/12mm, 74°/19mm, 76-79°/22mm, 95-98°/50mm, 177°/760mm, d<sub>4</sub><sup>20</sup> 1.357, n<sub>D</sub><sup>20</sup> 1.4539.** Purify the sulfonyl chloride by repeated distillation to remove HCl formed from hydrolysis. **It is a fuming, corrosive liquid, handle in a good fumehood.** It is

hydrolysed by aqueous N NaOH at room temperature and is best stored in aliquots in sealed ampules under N<sub>2</sub>. [Davies & Dick *J Chem Soc* 484 1932, Klamann & Drahowzal *Monatsh Chem* 83 463 1952, Saunders et al. *Biochem J* 36 372 1942, *Beilstein* 4 IV 34.]

**Ethanethiol (ethyl mercaptan)** [75-08-1] M 62.1, b 32.9°/704mm, d<sup>52</sup> 0.83147, pK<sup>25</sup> 10.61. Dissolve the thiol in aqueous 20% NaOH, extract it with a small amount of \*benzene and then steam distil until clear. After cooling, the alkaline solution is acidified slightly with 15% H<sub>2</sub>SO<sub>4</sub> and the thiol is distilled off, dried with CaSO<sub>4</sub>, CaCl<sub>2</sub> or 4A molecular sieves, and fractionally distilled under nitrogen [Ellis & Reid *J Am Chem Soc* 54 1674 1932]. It has a foul odour, work in an efficient fume cupboard. [*Beilstein* 1 IV 1390.]

**Ethanol** [64-17-5] M 46.1, b 78.3°, d<sup>15</sup> 0.79360, d<sup>5</sup> 0.78506, n<sub>D</sub><sup>20</sup> 1.36139, pK<sup>25</sup> 15.93. Usual impurities of fermentation alcohol are fusel oils (mainly higher alcohols, especially pentanols), aldehydes, esters, ketones and water. With synthetic alcohol, likely impurities are water, aldehydes, aliphatic esters, acetone and diethyl ether. Traces of \*benzene are present in ethanol that has been dehydrated by azeotropic distillation with \*benzene. Anhydrous ethanol is very *hygroscopic*. Water (down to 0.05%) can be detected by formation of a voluminous precipitate when aluminium ethoxide in \*benzene is added to a test portion. Rectified spirit (95% ethanol) is converted to *absolute* (99.5%) ethanol by refluxing with freshly ignited CaO (250g/L) for 6 hours, standing overnight and distilling with precautions to exclude moisture.

Numerous methods are available for further drying of *absolute* ethanol for making "Super dry ethanol". Lund and Bjerrum [*Chem Ber* 64 210 1931] used reaction with magnesium ethoxide, prepared by placing 5g of clean dry magnesium turnings and 0.5g of iodine (or a few drops of CCl<sub>4</sub>), to activate the Mg, in a 2L flask, followed by 50-75 ml of *absolute* ethanol, and warming the mixture until a vigorous reaction occurs. When this subsides, heating is continued until all the magnesium is converted to magnesium ethoxide. Up to 1L of ethanol is then added and, after refluxing for an hour, it is distilled off. The water content should be below 0.05%. Walden, Ulich and Laun [*Z Phys Chem* 114 275 1925] used **amalgamated aluminium** chips, prepared by degreasing aluminium chips (by washing with Et<sub>2</sub>O and drying in a vacuum to remove grease from machining the Al), treating with alkali until hydrogen evolved vigorously, washing with H<sub>2</sub>O until the washings were weakly alkaline and then stirring with 1% HgCl<sub>2</sub> solution. After 2 minutes, the chips were washed quickly with H<sub>2</sub>O, then alcohol, then ether, and dried with filter paper. (The amalgam became warm.) These chips were added to the ethanol, which was then gently warmed for several hours until evolution of hydrogen ceased. The alcohol was distilled and aspirated for some time with pure dry air. Smith [*J Chem Soc* 1288 1927] reacted 1L of *absolute* ethanol in a 2L flask with 7g of clean dry sodium, and added 25g of pure ethyl succinate (27g of pure ethyl phthalate was an alternative), and refluxed the mixture for 2 hours in a system protected from moisture, and then distilled the ethanol. A modification used 40g of ethyl formate instead, so that sodium formate separated out and, during reflux the excess of ethyl formate decomposed to CO and ethanol.

Drying agents suitable for use with ethanol include Linde type 4A molecular sieves, calcium metal, and CaH<sub>2</sub>. The calcium hydride (2g) is crushed to a powder and dissolved in 100ml of *absolute* ethanol by gently boiling. About 70ml of the ethanol are distilled off to remove any dissolved gases before the remainder is poured into 1L of *ca* 99.9% ethanol in a still, where it is boiled under reflux for 20 hours, while a slow stream of pure, dry hydrogen (better use nitrogen or Ar) is passed through. It is then distilled [Rüber *Z Elektrochem* 29 334 1923]. If calcium is used for drying, about ten times the theoretical amount should be used, and traces of ammonia (from some calcium nitride in the Ca metal) would be removed by passing dry air into the vapour during reflux. Ethanol can be freed from traces of basic materials by distillation from a little 2,4,6-trinitrobenzoic acid or sulfanilic acid. \*Benzene can be removed by fractional distillation after adding a little water (the \*benzene/water/ethanol azeotrope distils at 64.9°), the alcohol is then re-dried using one of the methods described above. *Alternatively*, careful fractional distillation can separate \*benzene as the \*benzene/ethanol azeotrope (b 68.2°). Aldehydes can be removed from ethanol by digesting with 8-10g of dissolved KOH and 5-10g of aluminium or zinc per L, followed by distillation. Another method is to heat under reflux with KOH (20g/L) and AgNO<sub>3</sub> (10g/L) or to add 2.5-3g of lead acetate in 5ml of water to 1L of ethanol, followed (slowly and without stirring) by 5g of KOH in 25ml of ethanol: after 1 hour the flask is shaken thoroughly, then set aside overnight before filtering and distilling. The residual water can be removed by standing the distillate over activated aluminium amalgam for 1 week, then filtering and distilling. Distillation of ethanol from Raney nickel

eliminates catalyst poisons.

Other purification procedures include pre-treatment with conc  $\text{H}_2\text{SO}_4$  (3ml/L) to eliminate amines, and with  $\text{KMnO}_4$  to oxidise aldehydes, followed by refluxing with  $\text{KOH}$  to resinify aldehydes, and distilling to remove traces of  $\text{H}_3\text{PO}_4$  and other acidic impurities after passage through silica gel, and drying over  $\text{CaSO}_4$ . Water can be removed by azeotropic distillation with dichloromethane (azeotrope boils at  $38.1^\circ$  and contains 1.8% water) or 2,2,4-trimethylpentane. [*Beilstein* 1 IV 1289.]

**Rapid purification:** Place degreased Mg turnings (grease from machining the turnings is removed by washing with dry EtOH then  $\text{Et}_2\text{O}$ , and drying in a vacuum) (5g) in a dry 2L round bottomed flask fitted with a reflux condenser (protect from air with a drying tube filled with  $\text{CaCl}_2$  or  $\text{KOH}$  pellets) and flushed with dry  $\text{N}_2$ . Then add iodine crystals (~0.5g) and gently warm the flask until iodine vapour is formed and coats the turnings. Cool, then add EtOH (50ml) and carefully heat to reflux until the iodine disappears. Cool again, then add more EtOH (to 1L) and reflux under  $\text{N}_2$  for several hours. Distil and store over 3A molecular sieves (pre-heated at  $300^\circ\text{--}350^\circ$  for several hours and cooled under dry  $\text{N}_2$  or argon).

**Ethoxycarbonyl isocyanate** [19617-43-7] **M 115.1, b 51-55°/13mm, 56°/18mm, 115-116°/781mm,  $d_4^{20}$  1.15.** Distil it twice from  $\text{P}_2\text{O}_5$  (1-2g) through a small Vigreux column and then through a 20-plate column. All fractional distillations should be under a vacuum. [*Lamon J Heterocycl Chem* 5 837 1968, *Beilstein* 3 H 36, 3 I 17.]

**Ethoxycarbonyl isothiocyanate** [16182-04-0] **M 131.5, b 43°/14mm, 51-55°/13mm, 56°/18mm,  $d_4^{20}$  1.12.** Fractionally distil it through a short column. It also distils at  $83^\circ/30\text{mm}$  with some decomposition liberating  $\text{CO}_2$  and sulfurous gases, best distil below 20mm vacuum. [*Capp et al. J Chem Soc* 1340, 1948, *Lamon J Heterocycl Chem* 5 837 1968, *Beilstein* 3 H 174, 3 I 71, 3 III 279, 3 IV 323.]

**2-Ethoxyethanol** [110-80-5] **M 90.1, b 134.8°,  $d_4^{20}$  0.931,  $n_D^{20}$  1.40751.** Dry it with  $\text{CaSO}_4$  or  $\text{K}_2\text{CO}_3$ , filter and fractionally distil it. Peroxides can be removed by refluxing with anhydrous  $\text{SnCl}_2$  or by filtration under slight pressure through a column of activated alumina. [*Beilstein* 1 IV 2377.]

**2-Ethoxyethyl methacrylate** [2370-63-0] **M 158.2, b 91-93°/35mm,  $d_4^{20}$  0.965,  $n_D^{20}$  1.429.** Purify the ester as described under methyl methacrylate below. [*Beilstein* 2 III 1291.]

**Ethyl acetate** [141-78-6] **M 88.1, b 77.1°,  $d_4^{20}$  0.9003,  $n_D^{20}$  1.37239,  $n_D^{25}$  1.36979,  $pK^{25}$  -6.93 (aqueous  $\text{H}_2\text{SO}_4$ ).** The most common impurities in EtOAc are water, EtOH and acetic acid. These can be removed by washing with aqueous 5%  $\text{Na}_2\text{CO}_3$ , then with saturated aqueous  $\text{CaCl}_2$  or  $\text{NaCl}$ , and drying with  $\text{K}_2\text{CO}_3$ ,  $\text{CaSO}_4$  or  $\text{MgSO}_4$ . More efficient drying is achieved if the solvent is further dried with  $\text{P}_2\text{O}_5$ ,  $\text{CaH}_2$  or molecular sieves before distillation.  $\text{CaO}$  has also been used. *Alternatively*, ethanol can be converted to ethyl acetate by refluxing with acetic anhydride (*ca* 1ml per 10ml of ester), the liquid is then fractionally distilled, dried with  $\text{K}_2\text{CO}_3$  and redistilled. [*Beilstein* 2 III 127.]

**Rapid purification:** Distil, dry over  $\text{K}_2\text{CO}_3$ , re-distil and store over 4A molecular sieves.

**Ethyl acetimidate** [1000-84-6] **M 87.1, b 92-95°/atm, 89.7-90°/765mm,  $d_4^{20}$  0.8671,  $n_D^{20}$  1.4025,  $pK_{\text{Est}} \sim 5.5$ .** It is best to prepare it freshly from the *hydrochloride* (see below). Dissolve the hydrochloride (123.5g) by adding it slowly to an ice-cold mixture of  $\text{H}_2\text{O}$  (500ml),  $\text{K}_2\text{CO}_3$  (276g) and  $\text{Et}_2\text{O}$  (200ml) and stirring rapidly. The  $\text{Et}_2\text{O}$  layer is separated, the aqueous layer is extracted with  $\text{Et}_2\text{O}$  (100ml), the combined  $\text{Et}_2\text{O}$  layers are dried ( $\text{MgSO}_4$ ), evaporated and the residual oil is distilled through a glass helices packed column (70x1.2cm). The yield is 19g (22%). [*Glickman & Cope J Am Chem Soc* 67 1020 1945, *Chaplin & Hunter J Chem Soc* 1118 1937, *Hunter & Ludwig Methods Enzymol* 25 585 1972.]

**Ethyl acetimidate hydrochloride** [2208-07-3] **M 123.6, m 98-100°(dec), 110-115°(dec), 112-113°(dec), m 112-114°(dec),  $pK_{\text{Est}} \sim 5.5$ .** Recrystallise the hydrochloride by dissolving it in the minimum volume of super dry EtOH and adding dry  $\text{Et}_2\text{O}$  or from dry  $\text{Et}_2\text{O}$ . Dry it in a vacuum and store it in a vacuum desiccator with  $\text{P}_2\text{O}_5$ . *Alternatively*, it could be crystallised from EtOH (containing a few of drops of ethanolic  $\text{HCl}$ ) and adding dry  $\text{Et}_2\text{O}$ . Filter and dry it in a vacuum desiccator over  $\text{H}_2\text{SO}_4$  and  $\text{NaOH}$ . [*Pinner Chem Ber* 16 1654 1883,

Glickman & Cope *J Am Chem Soc* **67** 1020 1945, Chaplin & Hunter *J Chem Soc* 1118 1937, McElvain & Schroeder *J Am Chem Soc* **71** 40 1949, McElvain & Tate *J Am Chem Soc* **73** 2233 1951, *Methods Enzymol* **25** 585 1972, *Beilstein* **2** III 418.]

**Ethyl acetoacetate** [141-97-9] **M 130.1, b 71°/12mm, 100°/80mm,  $d_4^{20}$  1.026,  $n_D^{20}$  1.419,  $pK^{25}$  10.68.** Shake the ester with small amounts of saturated aqueous  $\text{NaHCO}_3$  (until no further effervescence), then with water. Dry it with  $\text{MgSO}_4$  or  $\text{CaCl}_2$  and distil it under reduced pressure. [*Beilstein* **3** IV 1528.]

**Ethyl acrylate** [140-88-5] **M 100.1, b 20°/40mm, 99.5°/atm,  $d_4^{20}$  0.922,  $n_D^{20}$  1.406.** Wash the ester repeatedly with aqueous  $\text{NaOH}$  until free from inhibitors such as hydroquinone, then wash it with saturated aqueous  $\text{CaCl}_2$  and distil it under reduced pressure. Hydroquinone should be added if the ethyl acrylate is to be stored for extended periods. [*Beilstein* **2** IV 1460.] **LACHRYMATORY.**

**Ethylamine** [75-04-7] **M 45.1, b 16.6°/760mm,  $d_4^{20}$  1.3663,  $pK^{20}$  10.79.** Condense it in an all-glass apparatus cooled by circulating ice-water, and store it with  $\text{KOH}$  pellets below 0°. [*Beilstein* **4** IV 307.]

**Ethylamine hydrochloride** [557-66-4] **M 81.5, m 109-110°.** Crystallise the hydrochloride from absolute  $\text{EtOH}$  or  $\text{MeOH}/\text{CHCl}_3$ , wash it with dry ether and dry it in a vacuum. [*Beilstein* **4** IV 310.]

**Ethyl bromide** [74-96-4] **M 109.0, b 0°/165mm, 38°/745mm,  $d_4^{20}$  1.460,  $n_D^{20}$  1.4241.** The main impurities are usually  $\text{EtOH}$  and water, both of which form azeotropes with it. Ethanol and unsaturated compounds can be removed by washing with conc  $\text{H}_2\text{SO}_4$  until no further coloration is produced. The ethyl bromide is then washed with water, aqueous  $\text{Na}_2\text{CO}_3$ , and water again, then dried with  $\text{CaCl}_2$ ,  $\text{MgSO}_4$  or  $\text{CaH}_2$ , and distilled from  $\text{P}_2\text{O}_5$ . Olefinic impurities can also be removed by storing the ethyl bromide in daylight with elemental bromine, later removing the free bromine by extraction with dilute aqueous  $\text{Na}_2\text{SO}_3$ , drying the ethyl bromide with  $\text{CaCl}_2$  and fractionally distilling it. *Alternatively*, unsaturated compounds can be removed by bubbling oxygen containing ca 5% ozone through the liquid for an hour, then washing with aqueous  $\text{Na}_2\text{SO}_3$  to hydrolyse ozonides and remove hydrolysis products, followed by drying and distillation. [*Beilstein* **1** IV 150.]

**Ethyl bromoacetate** [105-36-2] **M 167.0, b 158-158.5°/758mm,  $d_4^{20}$  1.50,  $n_D^{20}$  1.450.** Wash the ester with saturated aqueous  $\text{Na}_2\text{CO}_3$  (three times), 50% aqueous  $\text{CaCl}_2$  (three times) and saturated aqueous  $\text{NaCl}$  (twice). Dry with  $\text{MgSO}_4$ ,  $\text{CaCl}_2$  or  $\text{CaCO}_3$ , and distil it. [*Beilstein* **2** IV 527.] **LACHRYMATORY.**

**Ethyl 6-bromohexanoate (ethyl 6-bromocaproate)** [25542-62-5] **M 223.1, b 120-125°/14mm, 126-127°/20-21mm, 128-130°/16mm, 127-130°/19mm,  $d_{23}^{23}$  1.241,  $d^{25}$  1.254,  $n_D^{20}$  1.458,  $n_D^{21}$  1.4566.** 6-Bromohexanoic acid (250g) is esterified by refluxing with  $\text{EtOH}$  (600ml) containing  $\text{H}_2\text{SO}_4$  (15ml) for 8 hours, evaporating *in vacuo*, the residue is taken up in  $\text{Et}_2\text{O}$ , washed with  $\text{H}_2\text{O}$ , then 5% of aqueous  $\text{Na}_2\text{CO}_3$  (effervescence), the  $\text{Et}_2\text{O}$  layer is dried ( $\text{CaCl}_2$ ), filtered, evaporated, and the residual oil is distilled to give the ester (178-218g). [Brown & Partridge *J Am Chem Soc* **66** 819 1844.] It has also been prepared from 6-hydroxyhexanoic acid by reaction with  $\text{PBr}_3$  in pyridine, or  $\text{HBr}$  ( $d$  1.5) and  $\text{H}_2\text{SO}_4$  and then esterifying in the same way [Barger et al. *J Chem Soc* 718 1937]. [*Beilstein* **2** IV 940.]

**Ethyl 2-(bromomethyl)acrylate** [17435-72-2] **M 193.1, b 38°/0.8mm,  $d_4^{20}$  1.398,  $n_D^{20}$  1.479.** If it contains some free acid, add  $\text{H}_2\text{O}$ , cool, and neutralise with  $\text{NaHCO}_3$  until evolution of  $\text{CO}_2$  ceases. Extract the mixture with  $\text{Et}_2\text{O}$  (3x) and dry the combined extracts ( $\text{Na}_2\text{SO}_4$ , 3 hours). Evaporate  $\text{Et}_2\text{O}$  and distil the ester collecting fraction **b** 39-40°/0.9mm, and check spectra. [Preparation and NMR: Ramarajan et al. *Org Synth Coll Vol VII* 211 1990, *Beilstein* **2** IV 1541.]

**Ethyl  $\alpha$ -bromopropionate** [535-11-5] **M 181.0, b 69-70°/25mm,  $d_4^{20}$  1.39,  $n_D^{20}$  1.447.** Wash the ester with saturated aqueous  $\text{Na}_2\text{CO}_3$  (three times), 50% aqueous  $\text{CaCl}_2$  (three times) and saturated aqueous  $\text{NaCl}$  (twice). Dry with  $\text{MgSO}_4$ ,  $\text{CaCl}_2$  or  $\text{CaCO}_3$ , and distil it. [*Beilstein* **2** IV 762.] **LACHRYMATORY.**

**Ethyl bromopyruvate** [70-23-5] **M 195.0, b 47°/0.5mm, 71-73°/5mm, 87°/9mm, 89-104°/14mm,  $d_4^{20}$  1.561,**

$n_D^{20}$  1.464. The most likely impurity is free acid (bromopyruvic or bromoacetic acids). Dissolve the ester in dry Et<sub>2</sub>O or dry CHCl<sub>3</sub>, stir with CaCO<sub>3</sub> until effervescence ceases, filter (may wash with a little H<sub>2</sub>O rapidly), dry (MgSO<sub>4</sub>) and distil it at least twice. The 2,4-dinitrophenylhydrazone has *m* 144-145°. [Burros & Holland *J Chem Soc* 672 1947, Letsinger & Laco *J Org Chem* 21 764 1956, Kruse et al. *J Am Chem Soc* 76 5796 1954, *Beilstein* 3 IV 1519.] **LACHRYMATORY.**

**2-Ethyl-1-butanol** [97-95-0] *M* 102.2, *b* 146.3°,  $n_D^{15}$  1.4243,  $n_D^{25}$  1.4205. Dry it with CaSO<sub>4</sub> for several days, filter and fractionally distil it. [*Beilstein* 1 IV 1725.]

**2-Ethylbut-1-ene** [760-21-4] *M* 84.1, *b* 66.6°,  $d_4^{20}$  0.833,  $n_D^{20}$  1.423. Wash it with 10N aqueous NaOH, then water. Dry the organic layer with CaCl<sub>2</sub>, filter and fractionally distil it. [*Beilstein* 1 IV 850.]

**Ethyl *n*-butyrate** [105-54-4] *M* 116.2, *b* 49°/50mm, 119-120°/760mm,  $d_4^{20}$  0.880,  $n_D^{20}$  1.393. Dry the ester with anhydrous CuSO<sub>4</sub> and distil it under dry nitrogen. [*Beilstein* 2 IV 787.]

**Ethyl carbamate** see **urethane** below.

**Ethyl carbazate (*N*-ethoxycarbonyl hydrazine)** [4114-31-2] *M* 104.1, *m* 44-48°, 51-52°, *b* 95.5°/10m, 92-95°/12mm, 100-102°/11mm. Fractionate the carbazate using a Vigreux column until the distillate crystallises [Allen & Bell *Org Synth Coll Vol III* 404 1955, *Beilstein* 3 IV 174].

**Ethyl chloride** [75-00-3] *M* 64.5, *b* 12.4°,  $d_4^{20}$  0.8978,  $n_D^{20}$  1.3676. Pass ethyl chloride through absorption towers containing, successively, conc H<sub>2</sub>SO<sub>4</sub>, NaOH pellets, P<sub>2</sub>O<sub>5</sub> on glass wool, or soda-lime, CaCl<sub>2</sub>, P<sub>2</sub>O<sub>5</sub>. Condense it into a flask containing CaH<sub>2</sub> and fractionally distil it. It has also been purified by illumination in the presence of bromine at 0° using a 1000W lamp, followed by washing, drying and distilling. [*Beilstein* 1 IV 124.]

**Ethyl chloroacetate** [105-39-5] *M* 122.6, *b* 143-143.2°,  $d_4^{20}$  1.150,  $n_D^{25}$  1.4192. Shake the ester with saturated aqueous Na<sub>2</sub>CO<sub>3</sub> (three times), aqueous 50% CaCl<sub>2</sub> (three times) and saturated aqueous NaCl (twice). Dry it with Na<sub>2</sub>SO<sub>4</sub> or MgSO<sub>4</sub> and distil it. [*Beilstein* 2 IV 481.] **LACHRYMATORY.**

**Ethyl chlorofluoroacetate (ClFCHCO<sub>2</sub>Et)** [ $\pm$  401-56-9] *M* 140.5, *b* 128°/atm, 133°/atm,  $d_4^{25}$  1.225,  $n_D^{25}$  1.3926. The ester is prepared from 2-chloro-1,1,2-trifluoroethyl ether (340g, 2.09mol at < 5°) while 96% H<sub>2</sub>SO<sub>4</sub> (228ml, 420g, 4.1mol) is added dropwise to it with stirring, and allowing the liberated HF (HIGHLY TOXIC vapour which etches glass) to vacate in an efficient fume cupboard, at a rate such that the temperature is maintained at 5-15° (ca 30-45 minutes). The mixture is stirred further at 10° for 2 hours when all the HF is released, then it is poured carefully onto crushed ice (1 Kg) and H<sub>2</sub>O (500ml). The lower white oily layer is allowed to settle, separated, washed until free of acid with saturated aqueous NaHCO<sub>3</sub> (3 x 25ml, **CARE**: CO<sub>2</sub> may be liberated), H<sub>2</sub>O (4 x 25ml) and dried over Drierite (10g) [Note that thorough washing is essential to avoid decomposition on distillation]. The crude dry ester is filtered, and fractionated through an efficient helix packed column (2 x 12cm) to give pure *ethyl chlorofluoroacetate* (190-100g, 65-68%) boiling at 129-130° and atmospheric pressure.

The *free acid* has *b* 162°/atm (160.5-161°/atm has also been reported),  $d_4^{25}$  1.532,  $n_D^{25}$  1.4085, is obtained by hydrolysis of the *methyl ester*, *b* 116°/atm, with 10% aqueous NaOH. The ethyl ester provides *chlorofluoroacetamide*, *b* 72-77°/1mm,  $d_4^{25}$  1.510,  $n_D^{25}$  1.4535, by reaction with ammonia [Young & Tarrant *J Am Chem Soc* 71 2432 1949, Hazeldine *J Chem Soc* 4259 1952]. [*Beilstein* 2 III 453, 2 IV 493, Englund *Org Synth Coll Vol IV* 423 1963, cf also Young & Terrant *J Am Chem Soc* 71 982 1949.]

**Ethyl chloroformate** [541-41-3] *M* 108.5, *m* -81°, *b* 94-95°,  $d_4^{20}$  1.135,  $n_D^{20}$  1.3974. Wash the ester several times with water, redistil it using an efficient fractionating column at atmospheric pressure and a CaCl<sub>2</sub> guard tube to keep free from moisture [Hamilton & Sly *J Am Chem Soc* 47 435 1925, Saunders et al. *J Am Chem Soc* 73 3796 1951]. [*Beilstein* 3 IV 23.] **LACHRYMATORY AND TOXIC.**

**Ethyl *trans*-crotonate** [623-70-1] *M* 114.2, *b* 137°,  $d_4^{20}$  0.917,  $n_D^{20}$  1.425. Wash it with aqueous 5% Na<sub>2</sub>CO<sub>3</sub>,

then with saturated aqueous  $\text{CaCl}_2$ , dry it with  $\text{CaCl}_2$  and distil it. [*Beilstein* 2 IV 1500.]

**Ethyl cyanoacetate** [105-56-6] M 113.1, b 206.0°,  $d_4^{20}$  1.061,  $n_D^{20}$  1.41751. Shake the ester several times with aqueous 10%  $\text{Na}_2\text{CO}_3$ , wash it well with water, dry with  $\text{Na}_2\text{SO}_4$  and fractionally distil it. [*Beilstein* 2 IV 1889.]

**Ethyl cyanofornate** [623-49-4] M 99.1, b 113-114°/740mm, 116.5-116.8°/765.5mm,  $d_4^{20}$  1.0112,  $n_D^{20}$  1.3818. Dissolve the cyanofornate in  $\text{Et}_2\text{O}$ , dry it over  $\text{Na}_2\text{SO}_4$ , filter, evaporate and distil it [Malachowsky et al. *Chem Ber* 70 1016 1937, Adickes et al. *J Prakt Chem* [2] 133 313 1932, Grundmann et al. *Justus Liebigs Ann Chem* 577 77 1952]. [*Beilstein* 2 IV 1862.]

**Ethyl diazoacetate** [623-73-4] M 114.1, m -22°, b 42°/5mm, 45°/12mm, 85-86°/88mm, 140-141°/720mm, 140-143°/atm,  $d_4^{17.6}$  1.0852,  $n_D^{17.6}$  1.4588. It is a very volatile yellow oil with a strong pungent odour. **EXPLOSIVE [distillation even under reduced pressure is dangerous and may result in an explosion — TAKE ALL THE NECESSARY PRECAUTIONS IF DISTILLATION IS TO BE CARRIED OUT]**. It explodes in contact with conc  $\text{H}_2\text{SO}_4$ -trace acid causes rapid decomposition. It is slightly soluble in  $\text{H}_2\text{O}$ , but is miscible with  $\text{EtOH}$ ,  $^*\text{C}_6\text{H}_6$ , petroleum ether and  $\text{Et}_2\text{O}$ . To purify, dissolve it in  $\text{Et}_2\text{O}$  [using  $\text{CH}_2\text{Cl}_2$  instead of  $\text{Et}_2\text{O}$ , protects the ester from acid], wash it with 10% aqueous  $\text{Na}_2\text{CO}_3$ , dry ( $\text{MgSO}_4$ ), filter and repeat as many times as possible until the  $\text{Et}_2\text{O}$  layer loses its yellow colour, then remove the solvent below 20° (vacuum). Note that prolonged heating may lead to rapid decomposition and low yields. It can also be purified by steam distillation under reduced pressure but with considerable loss in yield. Place the residual oil in a brown bottle, keep below 10°, and use as soon as possible without distilling. For preparing esters usually the ethereal solution is used directly without purification. [Womack & Nelson *Org Synth Coll Vol III* 392 1955, UV: Miller & White *J Am Chem Soc* 79 5974 1957, Fieser 1 367 1967, *Beilstein* 3 IV 1495.]

**Ethyl dibromoacetate** [617-33-4] M 245.9, b 81-82°/14.5mm, 194°/atm,  $d^{22}$  1.9081,  $n^{22}$  1.4973. Wash the ester briefly with conc aqueous  $\text{NaHCO}_3$ , then with aqueous  $\text{CaCl}_2$ . Dry it with  $\text{CaSO}_4$  and distil it under reduced pressure. [Hornyak & Amis *J Am Chem Soc* 79 2079 1957, *Beilstein* 2 H 219, 2 I 97, 2 III 484, 2 IV 533.]

**Ethyl dichloroacetate** [535-15-9] M 157.0, b 54-55°/11mm, 131.0-131.5°/40mm,  $d_4^{20}$  1.28,  $n_D^{20}$  1.438. Shake the ester with aqueous 3%  $\text{NaHCO}_3$  to remove free acid, wash with distilled water, dry for 3 days with  $\text{CaSO}_4$  and distil it under reduced pressure. [*Beilstein* 2 IV 501.]

**Ethyl 3,3-diethoxypropionate** [10601-80-6] M 190.2, b 58.5°/1.5mm, 65°/2mm, 95-96°/12mm,  $d_4^{20}$  0.78,  $n_D^{25}$  1.4101. Dissolve it in dry  $\text{Et}_2\text{O}$ , and dry with solid  $\text{NaHCO}_3$ , filter, distil and carefully fractionate it [Dyer & Johnson *J Am Chem Soc* 56 223 1934]. [*Beilstein* 3 II 411.]

**Ethylene (ethene)** [74-85-1] M 28.0, m -169.4°, b -102°/700mm. Purify ethylene by passage through a series of towers containing molecular sieves, or anhydrous  $\text{CaSO}_4$ , or cuprous ammonium solution, then conc  $\text{H}_2\text{SO}_4$ , followed by KOH pellets. *Alternatively*, it has been condensed in liquid nitrogen, with melting, freezing and pumping to remove air before passage through an activated charcoal trap, followed by a further condensation in liquid air. A sputtered sodium trap was used to remove oxygen. [*Beilstein* 1 IV 677.]

**Ethylenediamine (1,2-diaminoethane)** [107-15-3] M 60.1, f 11.0°, b 117.0°,  $d_4^{20}$  0.897,  $n_D^{20}$  1.45677,  $n^{30}$  1.4513,  $\text{pK}_1^{25}$  6.86,  $\text{pK}_2^{25}$  9.92. It forms a constant-boiling (b 118.5°, *monohydrate*, m 10°) mixture with water (23w/w%). [It is *hygroscopic* and miscible with water.] Recommended purification procedure [Asthana & Mukherjee in J.F. Coetzee (ed), *Purification of Solvents*, Pergamon Press, Oxford, 1982 cf p 53]: to 1L of ethylenediamine is added 70g of type 5A Linde molecular sieves and shaken for 12 hours. The liquid is decanted and shaken for a further 12 hours with a mixture of  $\text{CaO}$  (50g) and KOH (15g). The supernatant is fractionally distilled (at 20:1 reflux ratio) in contact with freshly activated molecular sieves. The fraction distilling at 117.2°/760mm is collected. Finally it is fractionally distilled from sodium metal. All distillations and storage of ethylenediamine should be carried out under nitrogen to prevent reaction with  $\text{CO}_2$  and water. The material containing 30% water is dried with solid NaOH (600g/L) and heated on a water bath for 10 hours.

Above 60°, separation into two phases takes place. The hot ethylenediamine layer is decanted off, refluxed with 40g of sodium for 2 hours and distilled [Putnam & Kobe *Trans Electrochem Soc* **74** 609 1938]. Ethylenediamine is usually distilled under nitrogen. *Alternatively*, it is dried over type 5A Linde molecular sieves (70g/L), then a mixture of 50g of CaO and 15g of KOH/L, with further dehydration of the supernatant with molecular sieves followed by distillation *from* molecular sieves and, finally, from sodium metal. A spectroscopically improved material is obtained by shaking with freshly baked alumina (20g/L) before distillation. [Beilstein **4** IV 1166.]

***N,N'*-Ethylenediaminediacetic acid (EDDA)** [5657-17-0] **M 176.2, m 222-224°(dec), pK<sub>1</sub><sup>25</sup> 6.48, pK<sub>2</sub><sup>25</sup> 9.57 (for NH groups).** Crystallise EDDA from H<sub>2</sub>O. [Beilstein **4** IV 2446.]

**Ethylenediamine dihydrochloride** [333-18-6] **M 133.0, m >300°, pK<sub>1</sub><sup>25</sup> 6.86, pK<sub>2</sub><sup>25</sup> 9.92.** Crystallise the salt from H<sub>2</sub>O or H<sub>2</sub>O/EtOH. Wash the crystals with EtOH and dry them *in vacuo*. It sublimes on heating. [Beilstein **4** IV 1168.]

**Ethylenediaminetetraacetic acid (EDTA)** [60-00-4] **M 292.3, m 253°(dec), pK<sub>1</sub><sup>25</sup> 0.26, pK<sub>2</sub><sup>25</sup> 0.96, pK<sub>3</sub><sup>25</sup> 2.60, pK<sub>4</sub><sup>25</sup> 2.67, pK<sub>5</sub><sup>25</sup> 6.16, pK<sub>6</sub><sup>25</sup> 10.26.** Dissolve EDTA in aqueous KOH or ammonium hydroxide, and precipitate it twice with dilute HCl or HNO<sub>3</sub>. Boil it twice with distilled water to remove mineral acid, then recrystallise it from water or dimethylformamide. Dry it at 110°. It also recrystallises from boiling 1N HCl; wash the crystals with distilled H<sub>2</sub>O and dry them *in vacuo*. [Ma & Ray *Biochemistry* **19** 751 1980, Beilstein **4** IV 2449.]

**Ethylene dimethacrylate (ethylene glycol dimethacrylate)** [97-90-5] **M 198.2, b 98-100°/5mm, d<sub>4</sub><sup>20</sup> 1.053, n<sub>D</sub><sup>20</sup> 1.456.** Distil it through a short Vigreux column (p 11) at about 1mm pressure, in the presence of 3% (w/w) of phenyl-β-naphthylamine. [Beilstein **2** IV 1532.]

**Ethylene dimyristate [1,2-bis(myristoyloxy)ethane]** [627-84-9] **M 482.8, m 61.7°.** Crystallise the ester from \*benzene/MeOH or diethyl ether/MeOH, and dry it in a vacuum desiccator. It forms an inclusion compound with 25.9 mols of urea. [McGreer et al. *J Am Chem Soc* **74** 3441 1952, Beilstein **2** H 366, **2** II 327, **2** III 924, **2** IV 1133.]

**Ethylene dipalmitate [1,2-bis(palmitoyloxy)ethane]** [624-03-3] **M 538.9, m 69.1°, 71.2°.** Crystallise the ester from \*benzene/MeOH, diethyl ether/MeOH or Me<sub>2</sub>CO and dry it in a vacuum desiccator. It forms an inclusion compound with 28.2 mols of urea. [McGreer et al. *J Am Chem Soc* **74** 3541 1952, Beilstein **2** H 373, **2** I 166, **2** II 338, **2** III 926, **2** IV 1169.]

**Ethylene distearate [1,2-bis(stearoyloxy)ethane]** [627-83-8] **M 595.0, m 74.4-75°, 75.3°, 77°.** Crystallise the ester from \*benzene/MeOH, diethyl ether/MeOH or Me<sub>2</sub>CO and dry it in a vacuum desiccator. It forms an inclusion compound with 31 mols of urea. [McGreer et al. *J Am Chem Soc* **74** 3541 1952, Beilstein **2** H 380, **2** II 354, **2** III 1021, **2** IV 1223.]

**Ethylene glycol** [107-21-1] **M 62.1, b 68°/4mm, 197.9°/760mm, d<sub>4</sub><sup>20</sup> 1.0986, n<sub>D</sub><sup>15</sup> 1.43312, n<sub>D</sub><sup>25</sup> 1.43056, pK<sub>25</sub><sup>25</sup> 10.6.** It is very *hygroscopic*, and also likely to contain higher diols. Dry it with CaO, CaSO<sub>4</sub>, MgSO<sub>4</sub> or NaOH and distil it under vacuum. Dry further by reaction with a small amount of sodium under nitrogen (to remove moisture), reflux for several hours and distil. The distillate is then passed through a column of Linde type 4A molecular sieves and finally distil under nitrogen, from more molecular sieves. Then fractionally distil it. [Beilstein **1** IV 2369.]

**Ethylene glycol bis(β-aminoethylether)-*N,N'*-tetraacetic acid (EGTA)** [67-42-5] **M 380.4, m >245°(dec), pK<sub>1</sub><sup>20</sup> 1.15 (2.40), pK<sub>2</sub><sup>20</sup> 2.40 (2.50), pK<sub>3</sub><sup>20</sup> 8.40 (8.67), pK<sub>4</sub><sup>20</sup> 8.94 (9.22).** Dissolve EGTA in aqueous NaOH, precipitate it by adding aqueous HCl, wash it with water and dry at 100° *in vacuo*. [Beilstein **4** IV 217.]

**Ethylene glycol diacetate** [111-55-7] **M 146.2, b 190.1°, 79-81°/11mm, d<sub>25</sub><sup>25</sup> 1.4188, n<sub>D</sub><sup>20</sup> 1.4150.** Dry the diacetate with CaCl<sub>2</sub>, filter (excluding moisture) and fractionally distil it under reduced pressure. [Beilstein **2** IV

1541.]

**Ethylene glycol dibutyl ether** [112-48-1] M 174.3, b 78-80°/16mm, 200-201°/760mm,  $d_4^{20}$  1.105,  $n_D^{20}$  1.42. Shake the ether with aqueous 5% Na<sub>2</sub>CO<sub>3</sub>, dry with MgSO<sub>4</sub> and store it with chromatographic alumina to prevent peroxide formation. [Beilstein 1 III 2083, 1 IV 2382.]

**Ethylene glycol diethyl ether (1,2-diethoxyethane)** [629-14-1] M 118.2, m -74°, b 121.5°,  $d_4^{20}$  0.842,  $n_D^{20}$  1.392. After refluxing for 12 hours, a mixture of the ether (2L), conc HCl (27ml) and water (200ml) is added with slow passage of nitrogen. The solution is cooled, and KOH pellets are added slowly and with shaking until no more dissolves. The organic layer is decanted, treated with some KOH pellets and again decanted. It is then refluxed with, and distilled from sodium immediately before use. *Alternatively*, after removal of peroxides by treatment with activated alumina, the ether is refluxed in the presence of the blue ketyl formed by sodium-potassium alloy with benzophenone, then distilled. [Beilstein 1 H 468, 1 II 519, 1 III 2078, 1 IV 2379.]

**Ethyl formate** [109-94-4] M 74.1, b 54.2°,  $d_4^{20}$  0.921,  $d_4^{30}$  0.909,  $n_D^{20}$  1.35994,  $n_D^{25}$  1.3565. Free acid or alcohol is removed by standing the ester over anhydrous K<sub>2</sub>CO<sub>3</sub>, with occasional shaking, then decanting and distilling from P<sub>2</sub>O<sub>5</sub>. *Alternatively*, the ester can be kept over CaH<sub>2</sub> for several days, then distilled from fresh CaH<sub>2</sub>. It cannot be dried with CaCl<sub>2</sub> because it reacts rapidly with the ester to form a crystalline compound. [Beilstein 2 IV 23.]

**Ethyl iodide (iodoethane)** [75-03-6] M 156.0, b 72.4°,  $d_4^{20}$  1.933,  $n_D^{15}$  1.5682,  $n_D^{25}$  1.5104. Drying the iodide with P<sub>2</sub>O<sub>5</sub> is unsatisfactory, and with CaCl<sub>2</sub> it is incomplete. It is probably best to dry it with sodium wire and distil [Hammond et al. *J Am Chem Soc* 82 704 1960]. Exposure of ethyl iodide to light leads to rapid decomposition, with the liberation of iodine. Free iodine can be removed by shaking with several portions of dilute aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> (until the colour is discharged), followed by washing with water, drying (with CaCl<sub>2</sub>, then sodium), and distilling. The distilled ethyl iodide is stored, over mercury, in a dark bottle away from direct sunlight. Other purification procedures include passage through a 60cm column of silica gel, followed by distillation, and treatment with elemental bromine, extraction of free halogen with Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> solution, followed by washing with water, drying and distilling. Free iodine and HI have also been removed by direct distillation through a LeBel-Henninger column containing copper turnings. Purification by shaking with alkaline solutions, and storage over silver, are reported to be unsatisfactory. [Beilstein 1 IV 163.]

**Ethyl isobutyrate** [97-62-1] M 116.2, b 110°,  $d_4^{20}$  0.867,  $n_D^{20}$  1.388. Wash the ester with aqueous 5% Na<sub>2</sub>CO<sub>3</sub>, then with saturated aqueous CaCl<sub>2</sub>. Dry it over CaSO<sub>4</sub> and distil. [Beilstein 1 IV 846.]

**Ethyl isocyanate** [109-90-0] M 71.1, b 559.8°/759mm, 59-61°/760mm, 60-63°/~760mm,  $d_4^{20}$  0.9031,  $n_D^{20}$  1.3808. Fractionate the isocyanate through an efficient column preferably in an inert atmosphere and store it in aliquots in sealed tubes [Bieber *J Am Chem Soc* 74 4700 1952, Slocombe et al. *J Am Chem Soc* 72 1888 1950]. [Beilstein 4 IV 402.]

**Ethyl isovalerate** [108-64-5] M 130.2, b 134.7°,  $d_4^{20}$  0.8664,  $n_D^{20}$  1.39621,  $n_D^{25}$  1.3975. Wash the ester with aqueous 5% Na<sub>2</sub>CO<sub>3</sub>, then saturated aqueous CaCl<sub>2</sub>. Dry it over CaSO<sub>4</sub> and distil. [Beilstein 2 IV 898.]

**Ethyl levulinate (4-oxopentanoic acid ethyl ester)** [539-88-8] M 144.2, m 37.2°, b 106-108°/2mm, 138.8°/8mm, 203-205°/atm,  $d_4^{20}$  1.012,  $n_D^{20}$  1.423. Stir the ester with Na<sub>2</sub>CO<sub>3</sub> and charcoal, filter and distil. It is freely soluble in H<sub>2</sub>O and EtOH [IR, NMR: Sterk *Monatsh Chem* 99 1770 1968, Thomas & Schuette *J Am Chem Soc* 53 2328 1931, Cox & Dodds *J Am Chem Soc* 55 3392 1933]. [Beilstein 3 IV 1562.]

**Ethyl malonate monoamide** [7597-56-0] M 131.1, m 47-50°, 49.5-50°, 50°, b 130-135°/2mm. The amide crystallises from Et<sub>2</sub>O or by slow evaporation of an aqueous solution as colourless crystals [Snyder & Elston *J Am Chem Soc* 76 3039 1954, McAlvain & Schroeder *J Am Chem Soc* 71 45 1949, Rising et al. *J Biol Chem* 89 20 1930]. [Beilstein 2 IV 1887.]

**Ethyl methacrylate** [97-63-2] M 114.2, b 59°/100mm,  $d_4^{20}$  0.915,  $n_D^{20}$  1.515. Wash the ester successively

with 5% aqueous  $\text{NaNO}_2$ , 5%  $\text{NaHSO}_3$ , 5%  $\text{NaOH}$ , then water. Dry it over  $\text{MgSO}_4$ , add 0.2% (w/w) of phenyl- $\beta$ -naphthylamine, and distil it through a short Vigreux column [Schultz *J Am Chem Soc* **80** 1854 1958]. [Beilstein **2** IV 1523.]

**Ethyl methyl ether** [540-67-0] **M 60.1, b 7°/760mm, d<sup>0</sup> 0.725, n<sup>4</sup> 1.3420**. Dry the ether with  $\text{CaSO}_4$ , pass it through an alumina column (to remove peroxides), then fractionally distil it while collecting fractions in receivers kept below 0°. [Beilstein **1** H 314, **1** I 158, **1** II 311, **1** III 1288, **1** IV 1314.]

**3-Ethyl-2-methyl-2-pentene** [19780-67-7] **M 112.2, b 109°/757mm, 114.5°/760mm, d<sub>4</sub><sup>20</sup> 0.72468, n<sub>D</sub><sup>20</sup> 1.4124**. Purify it by preparative GLC on a column of 20% squalene on Chromosorb P at 70°. Alternatively, fractionate it under an inert atmosphere. It forms an azeotrope with methoxyethanol. [Beilstein **1** H 222, **1** III 8471, **1** IV 890.]

**Ethyl nitroacetate** [626-35-7] **M 133.1, b 42-43°/0.2mm, 71-72°/3mm, 93-96°/9mm, 194-195°/atm, d<sub>4</sub><sup>20</sup> 1.1953, n<sub>D</sub><sup>20</sup> 1.4260, pK<sup>25</sup> 5.82**. Purify the ester by repeated distillation. IR:  $\nu_{\text{max}}$  1748 ( $\text{CO}_2$ ), 1570 and 1337 ( $\text{NO}_2$ ), and  $800\text{cm}^{-1}$  [Haszeldine *J Chem Soc* 2525 1953]. The *hydrazine salt* crystallises from 95% EtOH or MeOH as yellow crystals **m** 104-105° [Ungnade & Kissinger *J Org Chem* **22** 1661 1957, Emmons & Freeman *J Am Chem Soc* **77** 4391 1955]. [Beilstein **2** IV 537.]

**Ethyl propionate** [105-37-3] **M 102.1, b 99.1°, d<sub>4</sub><sup>20</sup> 0.891, n<sup>15</sup> 1.38643, n<sub>D</sub><sup>20</sup> 1.38394**. Treat the ester with anhydrous  $\text{CuSO}_4$  and distil it under nitrogen. [Beilstein **2** IV 205.]

**Ethyl pyruvate** [617-35-6] **M 116.1, m -50°, b 44-45°/10mm, 56°/20mm, 69-71°/42mm, 63°/23mm, 155.5°/760mm, d<sub>4</sub><sup>20</sup> 1.047, n<sub>D</sub><sup>20</sup> 1.4052**. Shake the ester with 10ml portions of saturated aqueous  $\text{CaCl}_2$  solution (removes ethyl acetate) and the organic layer is removed by centrifugation, decantation and filtration, and is distilled under reduced pressure. Purification of small quantities is carried out *via* the bisulfite adduct: the ester (2.2ml) is shaken with saturated  $\text{NaHSO}_3$  (3.6ml), chilled in a freezing mixture when crystals separate rapidly (particularly if seeded). After 5 minutes EtOH (10ml) is added and the crystals are filtered off, washed with EtOH and  $\text{Et}_2\text{O}$  and dried. Yield *ca* 3g of *bisulfite adduct*. Then treat the adduct (16g) with saturated aqueous  $\text{MgSO}_4$  (32ml) and 40% formaldehyde (5ml) and shake, whereby the ester separates as an oil which is extracted with  $\text{Et}_2\text{O}$ . The extract is dried ( $\text{MgSO}_4$ ), filtered, evaporated and the residue is distilled (**b** 56°/20mm), and then redistilled (**b** 147.5°/750mm) to give 5.5g of pure ester. [Cornforth *Org Synth Coll Vol IV* 467 1963, Beilstein **3** IV 1513.]

**Ethyl stearate** [111-61-5] **M 312.5, m 33°, b 213-215°/15mm**. The solid portion is separated from the partially solid starting material, then crystallised twice from EtOH, dried by azeotropic distillation with \*benzene, and fractionally distilled through a spinning-band column at low pressure [Welsh *Trans Faraday Soc* **55** 52 1959]. [Beilstein **2** IV 1218.]

**Ethyl thiocyanate (ethyl rhodanide)** [542-90-5] **M 87.1, b 144-145°, d<sub>4</sub><sup>20</sup> 1.011, n<sub>D</sub><sup>20</sup> 1.462**. Fractionally distil the ester at atmospheric pressure. [Beilstein **2** IV 1218.] (**CARE LACHRYMATOR.**)

**Ethyl thioglycolate (ethyl 2-mercaptoacetate)** [623-51-8] **M 120.2, b 50-51°/10mm, 55°/17mm, 62.5-64°/22mm, 67-68°/24mm, 155-158°/atm, d<sub>4</sub><sup>20</sup> 1.096, n<sub>D</sub><sup>20</sup> 1.457**. Dissolve the thioglycolate in  $\text{Et}_2\text{O}$ , wash with  $\text{H}_2\text{O}$ , dry it over  $\text{Na}_2\text{SO}_4$ , filter, evaporate and distil the residue under reduced pressure [Bredereck et al. *Chem Ber* **90** 1837 1957]. The *Ni complex* [ $\text{Ni}(\text{SCH}_2\text{CO}_2\text{Et})_2$ ], when recrystallised twice from EtOH, gives crystals which become black when dried in a vacuum over  $\text{H}_2\text{SO}_4$ , **m** 104-105° [Dranet & Cefola *J Am Chem Soc* **76** 1975 1954]. [Beilstein **3** H 255.]

**N-Ethyl thiourea** [625-53-6] **M 104.2, m 110°**. Crystallise the thiourea from EtOH, MeOH or ether. [Beilstein **4** IV 374.]

**Ethyl trichloroacetate** [515-84-4] **M 191.4, b 100-100.5°/30mm, d<sub>4</sub><sup>20</sup> 1.383**. Shake the ester with saturated aqueous  $\text{Na}_2\text{CO}_3$  (three times), aqueous 50%  $\text{CaCl}_2$  (three times), saturated aqueous  $\text{NaCl}$  (twice), then distil it

over  $\text{CaCl}_2$ , and redistil it under reduced pressure. [*Beilstein* 2 IV 514.]

**Ethyl trifluoroacetate** [383-63-1] M 142.1, b 61.3°/750, 60-62°/atm, 62-64°/755mm,  $d_4^{20}$  1.191,  $n_D^{20}$  1.30738. It has been prepared by treating sodium trifluoroacetate with  $\text{Et}_2\text{SO}_4$ , or the acid with excess of EtOH and a small amount of  $\text{H}_2\text{SO}_4$ , distilling the azeotropic mixture of ester and EtOH, removing the latter with  $\text{CaCl}_2$ , filtering and fractionally distilling [Henne et al *J Am Chem Soc* 69 1819]. Fractionate it through a long Vigreux column. IR has  $\nu_{\text{max}}$  at 1800 ( $\text{CO}_2$ ) and 1000 (OCO)  $\text{cm}^{-1}$  [Fuson et al. *J Chem Phys* 20 1627 1952, Bergman *J Org Chem* 23 476 1958]. [*Beilstein* 2 II 186, 2 III 427, 2 IV 463.]

**Ethyl 4,4,4-trifluoroacetoacetate (ethyl 3-oxo-4,4,4-trifluorobutyrate)** [372-31-6] M 184.1, b 47-49°/25mm, 129-130°/atm, 131.8°/atm,  $d_4^{25}$  1.259,  $n_D^{20}$  1.375. The ester was prepared by the method of Swarts [*Bull Acad Roy Belg* 13 175 1927] in which ethyl trifluoroacetate (1mol) was added slowly to a stirred suspension of NaOEt (1mol) in EtOH (exothermic reaction), followed EtOAc (1mol), and refluxed overnight. An improved purification was by treatment with a concentrated aqueous solution of  $\text{NaHSO}_4$  (very small molar excess), followed by addition of a clear aqueous solution of  $\text{Cu}(\text{OAc})_2$ , and the organic layer was distilled off. The Cu derivative was filtered off and recrystallised from EtOH, washed with  $\text{Et}_2\text{O}$  to give green crystals of the copper chelate m 189° [UV has  $\lambda_{\text{max}}$  at 220nm ( $\epsilon$  11,000) and 270nm ( $\epsilon$  21,000), 3.0mg % in EtOH]. The dry Cu chelate was suspended in  $\text{Et}_2\text{O}$  (some dissolved) and  $\text{H}_2\text{S}$  was bubbled through until precipitation of CuS was complete. The CuS was filtered off, and the filtrate was carefully fractionation (to avoid losses) to provide the ester in 54% overall yield. Its UV has  $\lambda_{\text{max}}$  at 238nm ( $\epsilon$  3000) and 288nm ( $\epsilon$  400), in cyclohexane [Breslow et al. *J Am Chem Soc* 68 102, 1946, Henne et al. *J Am Chem Soc* 69 1819 1947, Haszeldine et al. *J Chem Soc* 609 1951, *Beilstein* 3 II 425, 3 III 1206, 3 IV 1548].

**Ethyl trifluoromethanesulfonate** [425-75-2] M 178.1, b 115°/atm, 118-120°/atm,  $d_4^{20}$  1.378,  $n_D^{20}$  1.336. The ester reacts slowly with  $\text{H}_2\text{O}$  and aqueous alkali. If its IR has no OH bands ( $\sim 3000 \text{ cm}^{-1}$ ) then purify it by redistillation. If OH bands are present, then dilute with dry  $\text{Et}_2\text{O}$  and shake (carefully) with aqueous  $\text{NaHCO}_3$  until effervescence ceases, then wash with  $\text{H}_2\text{O}$  and dry ( $\text{MgSO}_4$ ), filter, evaporate and distil the residue under a slight vacuum then at atmospheric pressure in a  $\text{N}_2$  atmosphere. **It is a powerful alkylating agent, and the fumes are very toxic — perform all operations in an efficient fume cupboard.** [Gramstad & Haszeldine *J Chem Soc* 173 1956, Howells & McCown *Chem Rev* 77 69 1977, *Beilstein* 3 IV 34.]

**S-Ethyl trifluorothioacetate** [383-64-2] M 158.1, b 88-90°/atm, 90.5°/760mm,  $d_4^{20}$  1.255,  $n_D^{20}$  1.372. If IR is free of OH bands then fractionally distil it; otherwise dilute the thio-ester with dry  $\text{Et}_2\text{O}$ , wash with 5% KOH and  $\text{H}_2\text{O}$ , dry over  $\text{MgSO}_4$  and fractionate it through an efficient column [Hauptschein et al. *J Am Chem Soc* 74 4005 1952]. [*Beilstein* 2 IV 567.] *Powerful obnoxious odour.*

**Ethyl vinyl ether** [109-92-2] M 72.1, b 35.5°,  $d_4^{20}$  0.755. It usually contains polymerisation inhibitors (usually amines, e.g. triethanolamine) which can be removed by fractional distillation. Redistil it from sodium. [*Beilstein* 1 IV 2049.] **LACHRYMATORY.**

**Fluoroacetamide** [640-19-7] M 77.1, m 108°. Crystallise fluoroacetamide from chloroform and dry it in a vacuum. **TOXIC** [*Beilstein* 2 IV 454.]

**Formaldehyde** [50-00-0] M 30.0, m -92°, b -79.6°/20mm, -19.5°/760mm,  $d_4^{20}$  0.815,  $pK^{25}$  13.27 (hydrate)\*. Technical aqueous formaldehyde (*formalin*) solution commonly contains added MeOH (8-10%) to inhibit oxidation to formic acid. As a rough guide the  $d_4^{18}$  vs [concentration in g of HCHO/100ml (100g) of aqueous solution] is as follows: 1.0054 [2.24 (2.23)], 1.0126 [4.66 (4.60)], 1.0311 [11.08 (10.74)], 1.0410 [14.15 (13.59)], 1.0568 [19.89 (18.82)], 1.0719 [25.44 (23.73)], 1.0853 [30.17 (27.80)], 1.1057 [37.72 (34.11)] and 1.1158 [41.87 (37.53)]; values in curved brackets are from alternative determinations. [Marvel & Porter *Org Synth Coll Vol I* 377 1941]. If pure formaldehyde is required, add KOH solution (1 mole KOH: 100 moles HCHO) to ~37% by weight of aqueous formaldehyde solution (*formalin*), or evaporate formalin to dryness, to

give **paraformaldehyde polymer**  $\{\text{HO}(\text{CH}_2\text{O})_n\text{H}$ , [30525-89-4] **m 120-170°** depending on  $n\}$  as a white solid which, after washing with water, is dried in a vacuum desiccator over  $\text{P}_2\text{O}_5$  or  $\text{H}_2\text{SO}_4$ . Formaldehyde is regenerated by heating the paraformaldehyde to 120° under vacuum, or by decomposing it with barium peroxide. The **monomer**, a colourless flammable gas, is passed through a glass-wool filter cooled to  $-48^\circ$  in a  $\text{CaCl}_2/\text{ice}$  mixture to remove particles of polymer, then dried by passage over  $\text{P}_2\text{O}_5$  and either condensed in a bulb immersed in liquid nitrogen or absorbed in ice-cold conductivity water. The gas or aqueous solutions have **pungent suffocating odours**, are LACHRYMATORY and SUSPECTED CARCINOGEN, handle carefully. Formalin is a disinfectant and a preservative of dead animal and plant tissues. [*Beilstein* 1 IV 3017.]

**Formaldehyde dimethyl acetal (dimethoxymethane, methylal, formal)** [109-87-5] **M 76.1, m  $-108^\circ$ , b  $41-42^\circ/736\text{mm}$ ,  $41-43^\circ/\text{atm}$ ,  $42-46^\circ/\text{atm}$ ,  $d_4^{20}$  0.8608,  $n_D^{20}$  1.35335.** It is a volatile flammable liquid which is soluble in three parts of  $\text{H}_2\text{O}$ , and is readily hydrolysed by acids. Purify it by shaking with an equal volume of 20% aqueous NaOH, stand for 20 minutes, dry over fused  $\text{CaCl}_2$ , filter and fractionally distil it through an efficient column. Store it over molecular sieves. [Buchler et al. *Org Synth Coll Vol III* 469 1955, Rambaud & Besserre *Bull Soc Chim Fr* 45 1955, IR: Wilmshurst *Can J Chem* 36 285 1958, *Beilstein* 1 IV 3026.]

**Formaldehyde dimethyl mercaptal (bis-[methylthio]methane)** [1618-26-4] **M 108.2, b  $44-47^\circ/13\text{mm}$ ,  $45.5^\circ/18\text{mm}$ ,  $148-149^\circ/760\text{mm}$ ,  $d_4^{20}$  1.0594,  $n_D^{20}$  1.5322.** **Work in an efficient fume cupboard as the substance may contain traces (or more) of methylmercaptan which has a very bad odour.** Dissolve the mercaptal in  $\text{Et}_2\text{O}$ , shake it with aqueous alkalis then dry it over anhydrous  $\text{K}_2\text{CO}_3$ , filter and distil it over  $\text{K}_2\text{CO}_3$  under a stream of  $\text{N}_2$ . If the odour is very strong, then allow all gas effluents to bubble through 5% aqueous NaOH solution which is then treated with dilute  $\text{KMnO}_4$  in order to oxidise MeSH to odourless products. Its UV has  $\lambda_{\text{max}}$  at 238 nm ( $\log \epsilon$  2.73) [Fehnel & Carmack *J Am Chem Soc* 71 90 1949, Fehér & Vogelbruch *Chem Ber* 91 996 1958, Böhme & Marz *Chem Ber* 74 1672 1941]. Oxidation with aqueous  $\text{KMnO}_4$  yields **bis-(methylsulfonyl)methane** which has **m 142-143°** [Fieccchi et al. *Tetrahedron Lett* 1681 1967]. [*Beilstein* 1 IV 3088.]

**Formamide** [75-12-7] **M 45.0, f  $2.6^\circ$ , b  $103^\circ/9\text{mm}$ ,  $210.5^\circ/760\text{mm}(\text{dec})$ ,  $d_4^{20}$  1.13,  $n_D^{20}$  1.44754,  $n_D^{25}$  1.44682.** Formamide is easily hydrolysed by acids and bases. It also reacts with peroxides, acid halides, acid anhydrides, esters and (on heating) alcohols, while strong dehydrating agents convert it to a nitrile. It is very **hygroscopic**. Commercial material often contains acids and ammonium formate. Vorhoek [*J Am Chem Soc* 58 2577 1956] added some bromothymol blue to formamide and then neutralise it with NaOH before heating to  $80-90^\circ$  under reduced pressure to distil off ammonia and water. The amide is again neutralised and the process is repeated until the liquid remained neutral on heating. Sodium formate is added, and the formamide is concentrated under reduced pressure at  $80-90^\circ$ . The distillate is again neutralised and redistilled. It is then fractionally crystallised in the absence of  $\text{CO}_2$  and water by partial freezing.

Formamide (specific conductance  $2 \times 10^{-7} \text{ ohm}^{-1} \text{ cm}^{-1}$ ) of low water content is dried by passage through a column of 3A molecular sieves, then deionised by treatment with a mixed-bed ion-exchange resin loaded with  $\text{H}^+$  and  $\text{HCONH}^-$  ions (using sodium formamide in formamide) [Notley & Spiro *J Chem Soc (B)* 362 1966]. [*Beilstein* 2 IV 45.]

**Formamidine acetate** [3473-63-0] **M 104.1, m  $159-161^\circ(\text{dec})$ ,  $164^\circ(\text{dec})$ ,  $\text{pK}_{\text{Est}} \sim 12$ .** Unlike the hydrochloride, the acetate salt is not hygroscopic. It is recrystallised from a small volume of acetic acid, by addition of EtOH, and the crystals are washed with EtOH then  $\text{Et}_2\text{O}$  and dried in a vacuum. [Taylor et al. *Org Synth* 46 39 1966, *Beilstein* 2 IV 82.]

**Formamidine sulfinic acid (thiourea-S-dioxide)** [1758-73-2] **M 108.1, m  $124-126^\circ(\text{dec})$ .** Dissolve it in five parts of aqueous 1:1%  $\text{NaHSO}_3$  at  $60-63^\circ$  (charcoal), then allow it to crystallise slowly, with agitation, at  $10^\circ$ . Filter and dry it immediately at  $60^\circ$  [Koniecki & Linch *Anal Chem* 30 1134 1958]. [*Beilstein* 3 I 36, 3 IV 145.]

**Formic acid** [64-18-6] **M 46.0 (anhydrous), f  $8.3^\circ$ , b  $25^\circ/40\text{mm}$ ,  $100.7^\circ/760\text{mm}$ ,  $d_4^{20}$  1.22,  $n$  1.37140,  $n_D^{25}$  1.36938,  $\text{pK}^{25}$  3.74.** Anhydrous formic acid can be obtained by direct fractional distillation under reduced pressure, the receiver being cooled in ice-water. The use of  $\text{P}_2\text{O}_5$  or  $\text{CaCl}_2$  as dehydrating agents is unsatisfactory. Reagent grade 88% formic acid can be dried satisfactorily by refluxing with phthalic anhydride

for 6 hours and then distilling it. *Alternatively*, if it is left in contact with freshly prepared anhydrous  $\text{CuSO}_4$  for several days about one half of the water is removed from 88% formic acid; distillation then removes the remainder. Boric anhydride (prepared by melting boric acid in an oven at a high temperature, cooling in a desiccator, and powdering) is a suitable dehydrating agent for 98% formic acid; after prolonged stirring with the anhydride the formic acid is distilled under vacuum. Formic acid can be further purified by fractional crystallisation using partial freezing. [*Beilstein* 2 IV 3.]

***N*-Formyl *tert*-butylamine (*N*-*tert*-butylformamide)** [2425-74-3] **M 101.2, m 16°, b 48°/0.2mm, 78-83°/9mm, 135-136°/107mm, 202°/760mm,  $d_4^{25}$  0.903,  $n_D^{20}$  1.4330.** If the IR indicates some hydrolysis, then dissolve it in  $\text{Et}_2\text{O}$ , wash it with 20% aqueous  $\text{Na}_2\text{CO}_3$ , dry it ( $\text{MgSO}_4$ ), filter and fractionate it. Collect the fraction that solidifies on cooling and recrystallise it from  $\text{Et}_2\text{O}$  at low temperature if necessary. [*Emmons J Am Chem Soc* 79 5753 1957, *Beilstein* 4 III 324, 4 IV 661.]

***N*-Formyl ethylamine (*N*-ethylformamide)** [627-45-2] **M 73.1, b 29°/0.5mm, 176-179°/758mm,  $d_4^{20}$  0.950,  $n_D^{20}$  1.4346.** If the IR is good, then distil it and collect the middle fraction and redistil if necessary; otherwise proceed as for the previous amide. [*Erickson J Org Chem* 20 1569 1955, *Beilstein* 4 H 109, 4 I 352, 4 II 601, 4 III 207, 4 IV 346.]

**Formyl hydrazine (formic acid hydrazide)** [624-84-0] **M 60.1, m 54°, 54-57°,  $pK_{\text{est}} \sim 2.5$ .** Recrystallise it from  $\text{EtOH}$  and dry it *in vacuo*. Store below 10°; it may disproportionate on storage to 1,2-diformyl hydrazine and hydrazine. It forms a blue  $[\text{Cu}(\text{CH}_4\text{N}_2\text{O})]\text{SO}_4$  salt with  $\text{CuSO}_4$ . [*Beilstein* 2 H 93, 2 III 127, 2 IV 85.]

**Formyloxy acetonitrile (cyanomethyl formate)** [150760-95-5] **M 85.1, b 62-64°/12mm, 172-173°/atm,  $d_4^{25}$  0.903,  $n_D^{20}$  1.4330.** Purify it by fractional distillation and redistilling the middle fraction. It is useful for the formylation of alcohols and amines. The  $^{13}\text{C}$  NMR has  $\delta$  ( $\text{CDCl}_3$ ) at 47.87, 114.47 and 159.46 ppm. [*Deutsch & Niclas Synth Commun* 23 1561 1993, *Duczek et al. Synthesis* 37 1966.]

**Fumaraldehyde bis-(dimethyl acetal) (*trans*-1,1,4,4-tetramethoxybut-2-ene)** [6068-62-8] **M 176.2, b 100-103°/15mm, 101-103°/25mm,  $d_4^{20}$  1.011,  $n_D^{20}$  1.425.** Dry it over fused  $\text{CaCl}_2$  and distil it *in vacuo*. The maleic (*cis*) isomer has **b 112°/11mm, and  $d_4^{23}$  0.932 and  $n_D^{25}$  1.4243.** [*Zeik & Heusner Chem Ber* 90 1869 1957, *Clauson-Kaas et al. Acta Chem Scand* 9 111 1955, *Clauson-Kaas Acta Chem Scand* 6 569 1952, *Beilstein* 1 IV 3754.]

**Fumaric (*trans*-but-2-ene-1,4-dioic) acid** [110-17-8] **M 116.1, m 289.5-291.5° (sealed tube),  $pK_1^{25}$  3.10,  $pK_2^{25}$  4.60 (4.38).** Crystallise it from hot  $\text{M HCl}$  or water and dry it at 100°. [*Beilstein* 2 IV 2202.]

**Geraniol (*trans*-3,7-dimethyl-2,6-octadien-8-ol)** [106-24-1] **M 154.3, b 114-115°/11-12mm, 230°,  $d_4^{20}$  0.879,  $n_D^{20}$  1.4766.** Purify geraniol by ascending chromatography or by thin layer chromatography on plates of kieselguhr G with acetone/water/liquid paraffin (130:70:1) as solvent system. Hexane/ethyl acetate (1:4) is also suitable. Also purify it by GLC on a silicone-treated column of Carbowax 20M (10%) on Chromosorb W (60-80 mesh). [*Porter Pure Appl Chem* 20 499 1969.] Store it in full, tightly sealed containers in the cool and protect from light. It has a pleasant odour. [cf p 681, *Beilstein* 1 IV 2277.]

**Glutaraldehyde** [111-30-8] **M 100.1, b 71°/10mm, as 50% aqueous solution.** Likely impurities are oxidation products-acids, semialdehydes and polymers. It can be purified by repeated washing with activated charcoal (Norit) followed by vacuum filtration, using 15-20g charcoal/100ml of glutaraldehyde solution. Distil it at 60-65°/15mm, discarding the first 5-10%, then dilute with an equal volume of freshly distilled water at 70-75°, using magnetic stirring under nitrogen. The solution is stored at low temperature (3-4°), in a tightly stoppered container, and protected from light. Standardise by titration with hydroxylamine. [*Anderson J Histochem Cytochem* 15 652 1967, *Beilstein* 1 IV 3659.]

**Glutaric acid** [110-94-1] **M 132.1, m 97.5-98°,  $pK_1^{25}$  4.35,  $pK_2^{25}$  5.40.** Crystallise the acid from \*benzene,  $\text{CHCl}_3$ , distilled water or \*benzene containing 10% (w/w) of diethyl ether. Dry it under vacuum. [*Beilstein* 2

IV 1934.]

**dl-Glyceraldehyde** [56-82-6] **M 90.1, m 145°**. Crystallise it from EtOH/diethyl ether. The D(+)-enantiomer [453-17-8] is a syrup (70 + % H<sub>2</sub>O) with  $[\alpha]_{\text{D}}^{25} +14^{\circ}$  (c 2, H<sub>2</sub>O) and the *dimethyl acetal* has **b** 124-127 °/14mm and  $[\alpha]_{\text{D}}^{15} +21^{\circ}$  (c 18, H<sub>2</sub>O). [Beilstein 1 H 845, 1 IV 4114.]

**Glycerol** [56-81-5] **M 92.1, m 18.2°, b 182°/20mm, 290°/760mm, d<sub>4</sub><sup>20</sup> 1.261, n<sub>D</sub><sup>25</sup> 1.47352, pK<sup>25</sup> 14.4**. Glycerol is dissolved in an equal volume of *n*-butanol (or *n*-propanol, amyl alcohol or liquid ammonia) in a water-tight container, cooled and seeded while slowly revolving in an ice-water slurry. The crystals are collected by centrifugation, then washed with cold acetone or isopropyl ether. [Hass & Patterson *Ind Eng Chem (Anal Ed)* 33 615 1941.] Coloured impurities can be removed from substantially dry glycerol by extraction with 2,2,4-trimethylpentane. *Alternatively*, glycerol can be decolorised and dried by treatment with activated charcoal and alumina, followed by filtering. Glycerol can be distilled at 15mm in a stream of dry nitrogen, and stored in a desiccator over P<sub>2</sub>O<sub>5</sub>. Crude glycerol can be purified by digestion with concentrated H<sub>2</sub>SO<sub>4</sub> and then saponification with a lime paste, re-acidification with H<sub>2</sub>SO<sub>4</sub>, filtration, treatment with an anion exchange resin and fractional distillation under vacuum. [Beilstein 1 IV 2751.]

**Glycolic (α-hydroxyacetic) acid** [79-14-1] **M 76.1, m 81°, pK<sup>25</sup> 3.62**. Crystallise it from diethyl ether. [Beilstein 3 IV 571.]

**Guanidine** [113-00-8] **M 59.1, m 47.5-48.5°, 48-49°, ~50°, pK<sup>25</sup> 13.6**. Crystallise it from water/EtOH under nitrogen. It is very deliquescent and absorbs CO<sub>2</sub> from the air readily. [Jones *Trans Faraday Soc* 55 524 1959, *Beilstein* 3 H 82, 3 I 39, 3 II 69, 3 III 154, 3 IV 148.]

**Guanidine carbonate** [593-85-1] **M 180.2, m 197°, 230°**. Crystallise it from MeOH. [Beilstein 3 H 86, 3 I 41, 3 II 72, 3 III 161, 3 IV 152.]

**Guanidine hydrochloride** [50-01-1] **M 95.5, m 181-183°**. Crystallise the hydrochloride from hot methanol by chilling to about -10°, with vigorous stirring. The fine crystals are filtered through fritted glass, washed with cold (-10°) methanol, dried at 50° under vacuum for 5 hours. (The product is purer than that obtained by crystallisation at room temperature from methanol by adding large amounts of diethyl ether.) [Kolthoff et al. *J Am Chem Soc* 79 5102 1957, *Beilstein* 3 H 86, 3 II 71, 3 III 160, 3 IV 150.]

**Heptadecanoic acid (margaric)** [506-12-7] **M 270.5, m 60-61°, b 227°/100mm, pK<sub>Est</sub> ~4.9**. Crystallise the acid from MeOH or petroleum ether. [Beilstein 2 IV 1193.]

**1-Heptadecanol** [1454-85-9] **M 256.5, m 54°**. Crystallise it from acetone. [Beilstein 1 IV 1884.]

**Heptafluoro-2-iodopropane** [677-69-0] **M 295.9, b 39°/735mm, 41°/760mm, d<sub>4</sub><sup>0</sup> 2.1306, n<sub>D</sub><sup>20</sup> 1.3281**. Purify it by gas chromatography on a triacetin (glyceryl triacetate) column, followed by bulb-to-bulb distillation at low temperature. Store it over Cu powder to stabilise it. UV has  $v_{\text{max}}$  at 271nm ( $\epsilon$  240) in petroleum ether (b 60-80°). [Haszeldine *J Chem Soc* 1767, 3761 1953, *Beilstein* 1 III 255, 1 IV 225.]

***n*-Heptaldehyde** [111-71-7] **M 114.2, b 40.5°/12mm, 152.8°/760mm, d<sub>4</sub><sup>20</sup> 0.819, n<sub>D</sub><sup>25</sup> 1.4130**. Dry *n*-heptaldehyde with CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub> and fractionally distil it under reduced pressure. More extensive purification is by precipitation as the bisulfite compound (formed by adding the aldehyde to saturated aqueous NaHSO<sub>3</sub>) which is filtered off and recrystallised from hot H<sub>2</sub>O. The crystals, after being filtered and washed well with H<sub>2</sub>O, are hydrolysed by adding 700ml of aqueous Na<sub>2</sub>CO<sub>3</sub> (12.5% w/w of anhydrous Na<sub>2</sub>CO<sub>3</sub>) per 100g of aldehyde. The aldehyde is then steam distilled off, separated, dried with CuSO<sub>4</sub> and distilled under reduced pressure in a slow stream of nitrogen. [McNesby & Davis *J Am Chem Soc* 76 2148 1954, *Beilstein* 1 H 695, 1 I 357, 1 II 750, 1 III 2844, 1 IV 3314.]

***n*-Heptaldoxime** [629-31-2] M 129.2, m 53-55°. Separate the *cis*(Z) and *trans*(E) oximes by liquid chromatography through a silica gel column and eluting with petroleum ether (b 40-65°)/EtOAc (50:10) at a flow rate of 2-3.4 ml/sec where the *trans*-isomer comes through first and is a liquid with  $n_D^{22}$  1.38512, followed by the *cis*-isomer which is a solid, and crystallises from 60% aqueous EtOH with m 55°. They are identified by TLC on 0.2mm silica gel G by eluting with \*C<sub>6</sub>H<sub>6</sub>/EtOAc (50/10) and visualising with I<sub>2</sub> vapour: the *trans*-isomer has R<sub>F</sub> 0.6 and the *cis*-isomer has R<sub>F</sub> 0.5 [Pejkovic-Tadic et al. *J Chromatography* 21 239 1966, Emmous & Pagano *J Am Chem Soc* 77 4557 1955]. [*Beilstein* 1 H 698, 1 I 358, 1 II 752, 1 III 2850.]

***n*-Heptane** [142-18-5] M 100.2, b 98.4°,  $d_4^{20}$  0.684,  $n_D^{20}$  1.38765,  $n_D^{25}$  1.38512. Pass it through a silica gel column which greatly reduces the ultraviolet absorption of *n*-heptane. (The silica gel is previously heated to 350° before use.) For more extensive purification, heptane is shaken with successive small portions of conc H<sub>2</sub>SO<sub>4</sub> until the lower (acid) layer remains colourless. The heptane is then washed successively with water, aqueous 10% Na<sub>2</sub>CO<sub>3</sub>, water (twice), and dried with CaSO<sub>4</sub>, MgSO<sub>4</sub> or CaCl<sub>2</sub>. It is distilled from sodium. *n*-Heptane can be distilled azeotropically with methanol, then the methanol is washed out with water and, after drying, the heptane is redistilled. Other purification procedures include passage through activated basic Al<sub>2</sub>O<sub>3</sub>, drying with CaH<sub>2</sub>, storage with sodium, and stirring with 0.5N KMnO<sub>4</sub> in 6N H<sub>2</sub>SO<sub>4</sub> for 12 hours after treatment with conc H<sub>2</sub>SO<sub>4</sub>. Carbonyl-containing impurities have been removed by percolation through a column of impregnated Celite made by dissolving 0.5g of 2,4-dinitrophenylhydrazine in 6ml of 85% H<sub>3</sub>PO<sub>4</sub> by grinding together, then adding 4ml of distilled water and 10g Celite. [Schwartz & Parks *Anal Chem* 33 1396 1961, *Beilstein* 1 IV 376.]

**Hept-1-ene** [592-76-7] M 98.2, b 93°/771mm,  $d_4^{20}$  0.698,  $n_D^{20}$  1.400. Distil hept-1-ene from sodium, then carefully fractionally distilling it using an 18-in gauze-packed column. It can also be purified by azeotropic distillation with EtOH. It usually contains the 2- and 3-isomers as impurities. These can be removed by gas chromatography using a Carbowax column at 70°. [*Beilstein* 1 IV 857.]

***n*-Heptyl alcohol (1-heptanol)** [111-70-6] M 116.2, b 175.6°, d 0.825,  $n_D^{20}$  1.425. Shake the alcohol with successive lots of alkaline KMnO<sub>4</sub> until the colour persists for 15 minutes, then dry it with K<sub>2</sub>CO<sub>3</sub> or CaO, and fractionally distil it. [*Beilstein* 1 IV 1731.]

***n*-Heptylamine** [111-68-2] M 115.2, b 155°,  $d_4^{20}$  0.775,  $n_D^{20}$  1.434, pK<sup>25</sup> 10.66. Dry it in over KOH pellets for 24 hours, then decant it and fractionally distil it. Store away from CO<sub>2</sub>. [*Beilstein* 4 IV 734.]

***n*-Heptyl bromide** [629-04-9] M 179.1, b 70.6°/19mm, 180°/760mm,  $d_4^{20}$  1.140,  $n_D^{20}$  1.45. Shake it with conc H<sub>2</sub>SO<sub>4</sub>, wash with water, dry it with K<sub>2</sub>CO<sub>3</sub>, and fractionally distil. [*Beilstein* 1 IV 391.]

**Hexachloro-1,3-butadiene (perchlorobutadiene)** [87-68-3] M 260.8, b 144.1°/100mm, 210-212°/760mm,  $d_4^{20}$  1.683,  $n_D^{20}$  1.5556. Wash the diene with four or five 1/10th volumes of MeOH (or until the yellow colour has been extracted), then stir it for 2 hours with H<sub>2</sub>SO<sub>4</sub>, wash it with distilled water until neutral and filter it through a column of P<sub>2</sub>O<sub>5</sub>. Distil it under reduced pressure through a packed column. [Rytner & Bauer *J Am Chem Soc* 82 298 1960, *Beilstein* 1 IV 998.]

**Hexachloroethane** [67-72-1] M 236.7, m 187°. Steam distil it, then crystallise it from 95% EtOH. Dry it in the dark under vacuum. [*Beilstein* 1 IV 148.]

**Hexacosane (C-26)** [630-01-3] M 366.7, m 56.4°, b 169°/0.05mm, 205°/1mm, 262°/15mm. Distil hexacosane under vacuum and recrystallise it from diethyl ether. [*Beilstein* 1 IV 583.]

**Hexacosanoic acid (cerotinic acid)** [506-46-7] M 396.7, m 86-87°, 88-89°, pK<sub>Est</sub> ~4.9. Crystallise the acid from EtOH, aqueous EtOH and petroleum ether/Me<sub>2</sub>CO. [*Beilstein* 2 IV 1310.]

***n*-Hexadecane (Cetane)** [544-76-3] M 226.5, m 18.2°, b 105°/0.1mm,  $d_4^{20}$  0.773,  $n_D^{20}$  1.4345,  $n_D^{25}$  1.4325. Pass cetane through a column of silica gel and distil it under vacuum in a column packed with Pyrex helices. Store it over silica gel. It also crystallises from acetone, or is fractionally crystallised by partial freezing.

[Beilstein 1 IV 537.]

**1,14-Hexadecanedioic acid (Thaspic acid)** [505-54-4] M 286.4, m 126°, pK<sub>Est(1)</sub> ~4.5, pK<sub>Est(2)</sub> ~5.5. Recrystallise thaspic acid from EtOH, ethyl acetate or \*C<sub>6</sub>H<sub>6</sub>. [Beilstein 2 IV 2162.]

**Hexadecanoic acid (palmitic acid)** [57-10-3] M 256.4, m 62-63°, b 215°/15mm, pK<sup>25</sup> 6.46 (50% aqueous EtOH), 5.0 (H<sub>2</sub>O). Purify palmitic acid by slow (overnight) recrystallisation from hexane. Some samples are also crystallised from acetone, EtOH or EtOAc. The crystals are kept in air to lose solvent, or are pumped dry of solvent on a vacuum line. [Iwahashi et al. *J Chem Soc, Faraday Trans 1* 81 973 1985, pK: White *J Am Chem Soc* 72 1858 1950, Beilstein 2 IV 1157.]

**1,5-Hexadiene** [592-42-7] M 82.2, b 59.6°, d<sub>4</sub><sup>20</sup> 0.694, n<sub>D</sub><sup>20</sup> 1.4039. Distil 1,5-hexadiene from NaBH<sub>4</sub>. [Beilstein 1 IV 1013.]

**Hexafluoroacetone** [684-16-2, 34202-69-2 (3H<sub>2</sub>O)] M 166.1, m -129°, (trihydrate m 18-21°), b -28°. Dehydrate hexafluoroacetone by passing the vapours over P<sub>2</sub>O<sub>5</sub>. Ethylene is removed by passing the dried vapours through a tube containing Pyrex glass wool moistened with conc H<sub>2</sub>SO<sub>4</sub>. Further purification is by low temperature distillation using Warde-Le Roy stills. Store it in the dark at -78°. [Holmes & Kutschke *Trans Faraday Soc* 58 333 1962, Beilstein 1 IV 3215.]

**Hexafluoroacetylacetone (1,1,1,5,5,5-hexafluoro-2,4-pentanedione, hfacac)** [1522-22-1] M 208.1, b 63-65°/atm, 68°/736mm, 70-70.2°/760mm, 68-71°/atm, d<sub>4</sub><sup>20</sup> 1.490, n<sub>D</sub><sup>20</sup> 1.333. Hfacac has been prepared from a mixture of Na wire (0.34mole) covered with dry Et<sub>2</sub>O and CF<sub>3</sub>CO<sub>2</sub>Et (0.27mole see [383-63-1]) at 0°, to which was added trifluoroacetone (0.27mol, see [421-50-1]) at -78°. The mixture was then set aside overnight at ~25°, and any trifluoroacetone which distilled out was made to condense back into the reaction mixture which was stirred at ~25° until all the Na dissolved [Haszeldine et al. *J Chem Soc* 609 1951, compare with Henne et al. *J Am Chem Soc* 69 1819 1947]. The reddish-brown mixture was treated with an excess of 2N H<sub>2</sub>SO<sub>4</sub> and extracted with ether. The organic layer was dried (Na<sub>2</sub>SO<sub>4</sub>), and distilled to give hfacac (20g, 37%). By using NaOEt instead of Na metal, Henne and coworkers obtained a 72% yield of hfacac, which gave the *copper chelate* as bright grass green crystals (see below) [Henne et al. *J Am Chem Soc* 69 1819 1947]. Hfacac (100g) is purified by shaking twice with 98% H<sub>2</sub>SO<sub>4</sub> (300ml, use protective clothing) until it is completely dispersed, and set aside overnight. The anhydrous hfacac is separated and distilled through a glass-helices packed column with fraction b 70.2-70.5°/atm (~20g, analytically pure) being retained for metal complex studies [Buckingham et al. *Aust J Chem* 20 281 1967]. The oil readily forms a solid white stable *covalent dihydrate* [CF<sub>3</sub>-C(OH)<sub>2</sub>-CH<sub>2</sub>-C(OH)<sub>2</sub>-CF<sub>3</sub>] which is caused by the strong electron-withdrawing effect of the fluorine atoms. The dihydrate has no UV absorption spectrum; compare with λ<sub>max</sub> (CHCl<sub>3</sub>) 273nm (ε 7,800) for the anhydrous diketone. The dihydrate decomposes at ~90°. The hydrate (10g) can also be dehydrated by heating with anhydrous CaSO<sub>4</sub> (Drierite, 30g) and distilling; the distillate is treated with more CaSO<sub>4</sub> and redistilled (see above). When the distillate is treated with aqueous NaOH and heated, the dihydrate crystallises on cooling. The *Cu (II) complex* is more easily prepared from the anhydrous than from the hydrated diketone which needs to be dehydrated in the reaction medium, and has m 135° (after sublimation or crystallisation from CCl<sub>4</sub>). Store hfacac as the dihydrate; but the anhydrous diketone should be kept in an anhydrous atmosphere. [Gilman et al. *J Am Chem Soc* 78 2790 1956, Belford et al. *J Inorg Nucl Chem* 2 11 1956, Beilstein 1 III 3123, 1 IV 3681.]

**Hexafluoroethane** [76-16-4] M 138.0, b -79°. Purify it for pyrolysis studies by passing through a copper vessel containing CoF<sub>3</sub> at ca 270°, and held for 3 hours in a bottle with a heated (1300°) platinum wire. It is then fractionally distilled. [Steunenberg & Cady *J Am Chem Soc* 74 4165 1962, Beilstein 1 IV 132.]

**Hexafluoropropene (hexafluoropropylene, perfluoropropene, F-propene)** [116-15-4] M 150.0, m -153°, -156.2°, b -28°/atm, -29.4°/atm, d<sub>4</sub><sup>40</sup> 1.583. F-Propene is a gas that is available commercially in brass cylinders and is very **corrosive**, **toxic**, attacks skin and tissue membranes. It is prepared by treating CFCl<sub>2</sub>CFCICF<sub>3</sub> with Zn in boiling EtOH (3 days), but is faster under pressure at 100°. The gas is best handled

in a vacuum line and should be scrubbed over caustic alkali in a brass tube, as any HF present in it will attack glass. [Henne & Waalkes *J Am Chem Soc* **68** 496 1946, Henne & Hinkamp *J Am Chem Soc* **67** 1194 1945, *Beilstein* **1** III 697, **1** IV 735.]

**1,1,1,3,3,3-Hexafluoropropan-2-ol (HFIP)** [920-66-1] **M 168.1**, **m**  $-4^{\circ}$ , **b**  $57-58^{\circ}/760\text{mm}$ , **d**<sub>4</sub><sup>20</sup> **1.4563**, **n**<sub>D</sub><sup>22</sup> **1.2750**, **pK**<sub>25</sub><sup>25</sup> **9.30**. Distil it from 3A molecular sieves, retaining the middle fraction. It has been prepared by reduction of *hexafluoroacetone* in tetrahydrofuran (THF). In this case hexafluoropropanol forms a stable 1:1 THF complex which distils at  $99-100^{\circ}/760\text{mm}$  (**n**<sub>D</sub><sup>25</sup> 1.3283). The complex is decomposed by mixing with 20% oleum and distilling in a vacuum; the distillate is redistilled to give pure hexafluoropropan-2-ol with **b**  $59^{\circ}/760\text{mm}$ . The <sup>1</sup>H NMR shows a doublet at 4.52ppm (*J*<sub>H,H</sub> = 2Hz). The *benzoyl* derivative, [10315-85-2] **M 272.1**, has **m**  $53.9^{\circ}$  after crystallisation from pentane at  $-50^{\circ}$ , and its IR has  $\nu_{\text{max}}$  at  $1760\text{cm}^{-1}$ . [Middleton & Lindsey *J Am Chem Soc* **86** 4948 1964, Urry et al. *J Org Chem* **32** 347 1967.] It has very high peptide solubilising properties, alone or with CH<sub>2</sub>Cl<sub>2</sub> [use as a solvent: Narita et al. *Bull Chem Soc Jpn* **61** 281 1988, *Biochemistry* **29** 2639 1990.] It is **CORROSIVE**, causes severe eye irritation.

**Hexamethylenediamine (1,6-diaminohexane)** [124-09-4] **M 116.2**, **m**  $42^{\circ}$ , **b**  $46-47^{\circ}/1\text{mm}$ , **84.9^{\circ}/9\text{mm}, **100^{\circ}/20\text{mm}, **204-205^{\circ}/760\text{mm}, **pK**<sub>1</sub><sup>25</sup> **10.24**, **pK**<sub>2</sub><sup>25</sup> **11.02**. Crystallise it in a stream of nitrogen. It sublimes in a vacuum. [*Beilstein* **4** IV 1320.]******

**Hexamethylenediamine dihydrochloride** [6055-52-3] **M 189.2**, **m**  $248^{\circ}$ . Crystallise the salt from water or EtOH. [*Beilstein* **4** IV 1320.]

**Hexamethylene glycol (1,6-hexanediol)** [629-11-8] **M 118.2**, **m**  $41.6^{\circ}$ ,  $43-45^{\circ}$ , **b**  $134^{\circ}/10\text{mm}$ ,  $250^{\circ}$ , **n**<sub>D</sub><sup>20</sup> **1.458**. Fractionally crystallise it from its melt or from water. Distil it *in vacuo*. [*Beilstein* **1** IV 2556.]

***n*-Hexane** [110-54-3] **M 86.2**, **b**  $68.7^{\circ}$ , **d**<sub>4</sub><sup>20</sup> **0.660**, **n** **1.37486**, **n**<sub>D</sub><sup>25</sup> **1.37226**. Purify as for *n*-heptane. Modifications include the use of chlorosulfonic acid or 35% fuming H<sub>2</sub>SO<sub>4</sub> instead of conc H<sub>2</sub>SO<sub>4</sub> in washing the alkane, and final drying and distilling from sodium hydride. Unsaturated impurities can be removed by shaking the hexane with nitrating acid (58% H<sub>2</sub>SO<sub>4</sub>, 25% conc HNO<sub>3</sub>, 17% water, or 50% HNO<sub>3</sub>, 50% H<sub>2</sub>SO<sub>4</sub>), then washing the hydrocarbon layer with conc H<sub>2</sub>SO<sub>4</sub>, followed by H<sub>2</sub>O, drying, and distilling over sodium or *n*-butyl lithium. It can also be purified by distillation under nitrogen from sodium benzophenone ketyl solubilised with tetraglyme. Also purify it by passage through a silica gel column followed by distillation [Kajii et al. *J Phys Chem* **91** 2791 1987]. It is a **FLAMMABLE** liquid and a possible nerve toxin. [*Beilstein* **1** IV 338.]

**Rapid purification:** Distil, discarding the first forerun and stored over 4A molecular sieves.

**(±)-1,2-Hexanediol** [6920-22-5] **M 118.2**, **b**  $96-98^{\circ}/1\text{mm}$ ,  $118.4-118.5^{\circ}/13\text{mm}$ ,  $214-215^{\circ}/760\text{mm}$ , **d**<sub>4</sub><sup>20</sup> **0.951**, **n**<sub>D</sub><sup>20</sup> **1.442**. Fractionally distil it, preferably in a vacuum. Alternatively, dissolve it in Et<sub>2</sub>O, dry with K<sub>2</sub>CO<sub>3</sub> then Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil it in a vacuum. The *bis-4-nitrobenzoyl* derivative has **m**  $101.5-102.5^{\circ}$ . [Rudloff *Can J Chem* **36** 486 1958, *Beilstein* **1** I 251, **1** III 2200, **1** IV 2554.]

**1-Hexene** [592-41-6] **M 84.2**, **b**  $63^{\circ}$ , **d** **0.674**, **n**<sub>D</sub><sup>20</sup> **1.388**. Purify it by stirring over Na/K alloy for at least 6 hours, then fractionally distil it from sodium under nitrogen. [*Beilstein* **1** IV 828.]

***cis*-2-Hexene** [7688-21-3] **M 84.2**, **b**  $68-70^{\circ}$ , **d**<sub>4</sub><sup>20</sup> **0.699**, **n**<sub>D</sub><sup>20</sup> **1.399**. Purify it as for 1-hexene above. [*Beilstein* **1** IV 833.]

***trans*-2-Hexene** [4050-45-7] **M 84.2**, **b**  $65-67^{\circ}$ , **n**<sub>D</sub><sup>20</sup> **1.390**. Purify it as for 1-hexene above. [*Beilstein* **1** IV 834.]

***trans*-3-Hexene** [13269-52-8] **M 84.2**, **b**  $67-69^{\circ}$ , **d**<sub>4</sub><sup>20</sup> **0.678**, **n**<sub>D</sub><sup>20</sup> **1.393**. Purify it as for 1-hexene above. [*Beilstein* **1** IV 837.]

***n*-Hexyl alcohol (1-hexanol)** [111-27-3] **M 102.2**, **b**  $157.5^{\circ}$ , **d**<sub>4</sub><sup>20</sup> **0.818**, **n**<sub>D</sub><sup>15</sup> **1.4198**, **n**<sub>D</sub><sup>25</sup> **1.4158**. The commer-

cial material usually contains other alcohols which are difficult to remove. A suitable method is to esterify with hydroxybenzoic acid, recrystallise the ester and saponify. [Olivier *Rec Trav Chim Pays Bas* **55** 1027 1936.] Drying agents include  $K_2CO_3$  and  $CaSO_4$ , followed by filtration and distillation. (Some decomposition to the olefin occurs when Al amalgam is used as drying agent at room temperature, even if the amalgam is removed prior to distillation.) If the alcohol is required anhydrous, the redistilled material can be refluxed with the appropriate alkyl phthalate or succinate, as described under *Ethanol*. [Beilstein **1** IV 1694.]

***n*-Hexylamine** [111-26-2] **M 101.2, b 131°**,  $d_4^{20}$  0.765,  $n_D^{20}$  1.419,  $pK^{25}$  10.64. Dry with, and fractionally distil the hexylamine from, KOH or  $CaH_2$ . Store away from  $CO_2$ . [Beilstein **4** IV 709.]

***n*-Hexyl bromide** [111-25-1] **M 165.1, b 87-88°/90mm, 155°/743mm**,  $d_4^{20}$  1.176,  $n_D^{20}$  1.448. Shake the bromide with  $H_2SO_4$ , wash with water, dry ( $K_2CO_3$ ) and fractionally distil. [Beilstein **1** IV 352.]

***n*-Hexyl methacrylate** [142-09-6] **M 154.2, b 65-66°/4mm, 88-88.5°/14mm**,  $d_4^{20}$  0.8849,  $n_D^{20}$  1.4320. Purify it as for *methyl methacrylate*. [IR: Hughes & Walton *J Am Chem Soc* **79** 3985 1957, Beilstein **2** III 1288, **2** IV 1527.]

**Hexyltrimethylammonium bromide** [2650-53-5] **M 224.3, m 186°**. Recrystallise it from acetone. It is extremely *hygroscopic*. [McDowell and Kraus *J Am Chem Soc* **73** 2170 1951, Beilstein **4** IV 710.]

**1-Hexyne** [693-02-7] **M 82.2, b 12.5°/75mm, 71°/760mm**,  $d_4^{20}$  0.7156,  $n_D^{20}$  1.3989. Distil it from  $NaBH_4$  to remove peroxides. Stand over sodium for 24 hours, then fractionally distil it under reduced pressure. Also dry it by repeated vacuum transfer into freshly activated 4A molecular sieves, followed by vacuum transfer onto Na/K alloy, and stirring for 1 hour before fractionally distilling. [Beilstein **1** IV 1006.]

**2-Hexyne** [764-35-2] **M 82.2, b 83.8°/760mm**,  $d_4^{20}$  0.73146,  $n_D^{20}$  1.41382. Purify as for 1-hexyne above. [Beilstein **1** IV 1009.]

**3-Hexyne** [928-49-4] **M 82.2, b 81°/760mm**,  $d_4^{20}$  0.7231,  $n_D^{20}$  1.4115. Purify as for 1-hexyne above. [Beilstein **1** IV 1009.]

**(±)-5-Hexyn-3-ol (4-hydroxy-1-hexyne)** [19780-84-8] **M 98.1, b 58-59°/25mm, 73-76°/60mm**,  $d_4^{20}$  0.8918,  $n_D^{20}$  1.4437. Purify the hexynol by fractionation in a vacuum. The *carbamoyl* derivative (prepared by reaction with  $COCl_2$ /toluene followed by  $NH_3$ ) is crystallised by dissolving in the minimum volume of toluene and adding excess of petroleum ether (b 40-60°), and has **m** 70-71°. [Länger et al. *Helv Chim Acta* **42** 2379 1959, Beilstein **1** IV 2235.]

**Hydrazine *N,N'*-dicarboxylic acid diamide** [110-21-4] **M 116.1, m 245-246°(dec), 248°**. Crystallise the diamide from water, wash the crystals with EtOH then  $Et_2O$  and dry in vacuum over  $P_2O_5$ . It does **not** decompose on drying at 110°/48 hours. Its solubility in  $H_2O$  is 1% at 0°. [Andrieth & Mohr *Inorg Synth* **IV** 26 1953, Beilstein **3** H 116, **3** I 56, **3** III 229.]

**3-Hydroxy-2-butanone (acetoin)** [513-86-0] **M 88.1, b 144-145°**, [**m 100-105° dimer**]. Wash acetoin with EtOH until colourless, then with diethyl ether or acetone to remove biacetyl. Dry it in air by suction and dry further in a vacuum desiccator. [Beilstein **1** IV 3991.]

**(±)- $\alpha$ -Hydroxy- $\gamma$ -butyrolactone** [19444-84-9, S(-)- 733-52-4] **M 102.1, b 84°/0.2mm, 133°/10mm**,  $d_4^{20}$  1.310,  $n_D^{20}$  1.4656. It has been purified by repeated fractionation and forms a colourless liquid. It has to be distilled at high vacuum; otherwise it will dehydrate. The *acetoxo* derivative has **b** 94°/0.2mm. The *S-enantiomer* has  $d_4^{20}$  1.24,  $n_D^{20}$  1.464,  $[\alpha]_D^{25}$  -82° (c 2, MeOH). [NMR: Daremon & Rambaud *Bull Soc Chim Fr* 294 1971, Schmitz et al. *Chem Ber* **108** 1010 1975, Beilstein **18/1** V 5.]

**12-Hydroxydodecanoic acid** [505-95-3] **M 216.3, m 86-88°**,  $pK_{Est}$  ~4.8. Crystallise the acid from toluene

[Sadowik et al. *J Am Chem Soc* **108** 7789 1986]. [*Beilstein* **3** III 658.]

***N*-[2-Hydroxyethyl]ethylenediamine [2-(2-aminoethylamino)ethanol]** [111-41-1] **M 104.1, b 91.2°/5mm, 238-240°/752mm,  $d_4^{20}$  1.030,  $n_D^{20}$  1.485,  $pK_1^{20}$  3.75,  $pK_2^{20}$  9.15.** Distil the amine twice through a Vigreux column. Redistil it from solid NaOH, then from CaH<sub>2</sub>. *Alternatively*, it can be converted to the dihydrochloride and recrystallised from water. It is then dried, mixed with excess of solid NaOH and the free base is distilled from the mixture. It is finally redistilled from CaH<sub>2</sub>. [Drinkard et al. *J Am Chem Soc* **82** 2992 1960, *Beilstein* **4** IV 1558.]

***N*-[2-Hydroxyethyl]ethylenediaminetriacetic acid (HEDTA)** [150-39-0] **M 278.3, m 212-214°(dec),  $pK_1^{20}$  2.51,  $pK_2^{20}$  5.31,  $pK_3^{20}$  9.86.** Crystallise HEDTA from warm H<sub>2</sub>O, after filtering, by addition of 95% EtOH and allowing to cool. The crystals, collected on a sintered-glass funnel, are washed three times with cold absolute EtOH, then again crystallised from H<sub>2</sub>O. After leaching with cold H<sub>2</sub>O, the crystals are dried at 100° under vacuum. [Spedding et al. *J Am Chem Soc* **78** 34 1956, *Beilstein* **4** IV 2449.]

***N*-Hydroxyethyliminodiacetic acid (HIMDA)**[93-62-9] **M 177.2, m 181°(dec),  $pK_1^{25}$  2.16,  $pK_2^{25}$  8.72,  $pK_3^{25}$  13.7 (OH).** Crystallise HIMDA from water. [*Beilstein* **4** IV 2432.]

**2-Hydroxyethylimino-tris(hydroxymethyl)methane (MONO-TRIS)** [7343-51-3] **M 165.2, m 91°,  $pK_{Est}$  ~9.8.** Crystallise it twice from EtOH. Dry it under vacuum at 25°.

**2-Hydroxyethyl methacrylate** [868-77-9] **M 130.1, b 67°/3.5mm,  $d_4^{20}$  1.071,  $n_D^{20}$  1.452.** Dissolve the ester in water and extract with *n*-heptane to remove ethylene glycol dimethacrylate (checked by gas-liquid chromatography and by NMR) and distil it twice under reduced pressure [Strop et al. *J Phys Chem* **80** 694 1976]. [*Beilstein* **2** IV 1530.]

***dl*-2-Hydroxy-2-methylbutyric acid** [3739-30-8] **M 118.1, m 72-73°,  $pK^{25}$  3.73.** Crystallise the acid from \*benzene, and sublime it at 90°. [*Beilstein* **3** H 324.] **IRRITANT.**

***dl*-2-Hydroxy-3-methylbutyric ( $\alpha$ -hydroxyisovaleric) acid** [600-37-3] **M 118.1, m 86°,  $pK_{Est}$  ~3.9.** Crystallise the acid from ether/pentane. [*Beilstein* **3** IV 618.] **IRRITANT.**

***R*- $\gamma$ -Hydroxymethyl- $\gamma$ -butyrolactone** [52813-63-5] **M 116.1, b 101-102°/0.048mm,  $d_4^{20}$  1.2238,  $n_D^{20}$  1.471,  $[\alpha]_{546}^{20}$  -38°,  $[\alpha]_D^{20}$  -33° (c 3, EtOH),  $[\alpha]_D^{30}$  -53.5° (c 3, EtOH).** Purify it by column chromatography on Silica gel 60 (Merck 70-230 mesh) and elute with 7% EtOH/73% CHCl<sub>3</sub>. Its IR (film) has  $\nu_{max}$  at 3400 (OH), 1765 (C=O) and 1180 (COC) cm<sup>-1</sup>. [Eguchi & Kakuta *Bull Chem Soc Jpn* **47** 1704 1974, IR and NMR: Ravid et al. *Tetrahedron* **34** 1449 1978, *Beilstein* **3** III 620.]

**3-Hydroxy-3-methylglutaric acid (Meglutol)** [503-49-1] **M 162.1, m 99-102°, 108-109°, 100°,  $pK_{Est(1)}$  ~4.0,  $pK_{Est(2)}$  ~5.0.** Recrystallise the acid from diethyl ether/hexane and dry it under a vacuum at 60° for 1 hour. [*Beilstein* **3** IV 1166.]

**4-Hydroxy-4-methyl-2-pentanone** [123-42-2] **M 116.2, b 166°,  $d_4^{20}$  0.932,  $n_D^{20}$  1.4235,  $n_D^{25}$  1.4213.** The pentanone loses water when heated. It can be dried with CaSO<sub>4</sub>, then fractionally distilled under reduced pressure. [*Beilstein* **1** IV 403.]

**2-Hydroxy-2-methylpropionic acid ( $\alpha$ -hydroxyisobutyric acid, 2-methylactic acid)** [594-61-6] **M 104.1, m 79°, b 114°/12mm, 212°/760mm,  $pK^{25}$  3.78.** Distil the acid in steam, crystallise it from Et<sub>2</sub>O or \*benzene, dry it under vacuum, or sublime it at 50°. [*Beilstein* **3**, **7** IV 782.]

**( $\pm$ )-2-Hydroxyoctanoic acid (2-hydroxycaprylic acid)** [617-73-2] **M 160.2, m 69.5°, b 160-165°/10mm,  $pK_{Est}$  ~3.7.** Crystallise the acid from EtOH/petroleum ether or ether/ligroin. [*Beilstein* **3** IV 873.]

***N*-Hydroxysuccinimide** [6066-82-6] **M 115.1, m 96-98°,  $pK^{25}$  6.0.** Recrystallise the imide from EtOH/ethyl

acetate [Manesis & Goodmen *J Org Chem* **52** 5331 1987]. [*Beilstein* **21/9** V 498.]

**(±)-2-Hydroxytetradecanoic acid (α-hydroxymyristic acid)** [2507-55-3] **M 244.4, m 81-82°, pK<sub>Est</sub> ~3.7.** Crystallise the acid from chloroform or twice from MeOH (**m 85.8-86.6°**) [Horn & Pretorius *J Chem Soc* 1463 1954, Chibnall et al. *Biochem J* **30** 1034 1963, *Beilstein* **3** H 361, **3** I 130, **3** II 246, **3** III 660, **3** IV 921].

**R-2-Hydroxytetradecanoic acid** [26632-17-7] **M 244.4, m 88-2-88.5°, [α]<sub>D</sub><sup>20</sup> -3.1° (CHCl<sub>3</sub>).** Crystallise the acid from chloroform or first from Me<sub>2</sub>CO, then hexane [Horn & Pretorius *J Chem Soc* 1463 1954, Horn et al. *J Chem Soc* 177 1954, *Beilstein* **3** III 660.]

**Hydroxyurea** See in “Inorganic Compounds”, Chapter 5.

**Iminodiacetic acid** [142-73-4] **M 133.1, m 225°(dec), pK<sub>1</sub><sup>25</sup> 2.50, pK<sub>2</sub><sup>25</sup> 9.40.** Recrystallise iminodiacetic acid from water and dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **4** IV 2428.]

**Iodoacetamide** [144-48-9] **M 185.0, m ca 143°(dec).** Crystallise it from water or CCl<sub>4</sub>. It is used for tagging proteins. [Gurd *Methods Enzymol* **25** 424 1972, *Beilstein* **2** IV 536.]

**Iodoacetic acid** [64-69-7] **M 160.6, m 78°, pK<sup>25</sup> 3.19.** Crystallise it from petroleum ether (b 60-80°) or CHCl<sub>3</sub>/CCl<sub>4</sub>. [*Beilstein* **2** IV 534.]

**2-Iodobutane (sec-butyl iodide)** [513-48-4] **M 184.0, b 120.0, d<sub>4</sub><sup>20</sup> 1.50, n<sub>D</sub><sup>25</sup> 1.4973.** Purify the iodide by shaking with conc H<sub>2</sub>SO<sub>4</sub>, then washing it with water, aqueous Na<sub>2</sub>SO<sub>3</sub> and again with water. Dry (MgSO<sub>4</sub>) and distil. *Alternatively*, pass it through a column of activated alumina before distillation, or treat with bromine, followed by extraction of the free halogen with aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, thoroughly washing with water, drying and distilling. It is stored over silver powder and distilled before use. [*Beilstein* **1** IV 272.]

**Iodoform** [75-47-8] **M 393.7, m 119°.** Crystallise it from MeOH, EtOH or EtOH/EtOAc. It is steam volatile. It is a disinfectant. [*Beilstein* **1** IV 97.]

**N-Iodosuccinimide** [516-12-1] **M 225.0, m 200-201°.** Crystallise it from dioxane/CCl<sub>4</sub>. It iodates arenes in triflic acid. [Olah et al *J Org Chem* **58** 3194 1993, *Beilstein* **21/9** V 544.]

**Isoamyl acetate (1-butyl-3-methyl acetate, isopentyl acetate)** [123-92-2] **M 130.2, b 142.0°, d<sub>4</sub><sup>20</sup> 0.871, n<sub>D</sub><sup>20</sup> 1.40535.** Dry the acetate with finely divided K<sub>2</sub>CO<sub>3</sub> and fractionally distil it. [*Beilstein* **2** IV 157.]

**Isoamyl alcohol (3-methyl-1-butanol, 1-butyl-3-methyl alcohol)** [123-51-3] **M 88.2, b 128°/750mm, 132°/760mm, d<sub>15</sub><sup>15</sup> 0.8129, n<sub>15</sub><sup>15</sup> 1.4085, n<sub>D</sub><sup>20</sup> 1.4075.** Dry the alcohol by heating with CaO and fractionally distilling, then heating with BaO and redistilling. *Alternatively*, boil it with concentrated KOH solution, wash it with dilute H<sub>3</sub>PO<sub>4</sub>, and dry it with K<sub>2</sub>CO<sub>3</sub>, then anhydrous CuSO<sub>4</sub>, before fractionally distilling it. If very dry alcohol is required, the distillate is refluxed with the appropriate alkyl phthalate or succinate as described for *ethanol*. It is separated from 2-methyl-1-butanol by fractional distillation, fractional crystallisation and preparative gas chromatography. [*Beilstein* **1** IV 1677.]

**Isoamyl bromide (1-butyl-3-methyl bromide)** [107-82-4] **M 151.1, f -112°, b 119.2°/ 737mm, d<sub>4</sub><sup>20</sup> 1.208, n 1.444.** Shake the bromide with conc H<sub>2</sub>SO<sub>4</sub>, wash with water, dry with K<sub>2</sub>CO<sub>3</sub> and fractionally distil it. [*Beilstein* **1** IV 378.]

**Isoamyl chloride (1-butyl-3-methyl chloride)** [107-84-6] **M 106.6, b 99°/734mm, d<sub>4</sub><sup>20</sup> 0.8704, n<sub>D</sub><sup>20</sup> 1.4084.** Shake the chloride vigorously with 95% H<sub>2</sub>SO<sub>4</sub> until the acid layer no longer becomes coloured during 12 hours. then wash it with water, saturated aqueous Na<sub>2</sub>CO<sub>3</sub>, and more water. Dry it with MgSO<sub>4</sub>, filter and fractionally distil it. *Alternatively*, a stream of oxygen containing 5% of ozone is passed through the chloride for a time, three times longer than is necessary to cause the first coloration of starch iodide paper by the exit gas.

Subsequent washing of the liquid with aqueous  $\text{NaHCO}_3$  hydrolyses the ozonides and removes organic acids. After drying and filtering, the isoamyl chloride is distilled. [Chien & Willard *J Am Chem Soc* **75** 6160 1953, *Beilstein* **1** IV 287.]

**Isoamyl ether [diisopentyl ether, di-(1-butyl-3-methyl) ether]** [544-01-4] **M 158.3, b 173.4°, d<sub>4</sub><sup>20</sup> 0.778, n<sub>D</sub><sup>20</sup> 1.40850.** This is a mixture of 2- and 3-methylbutyl ether. It is purified by refluxing with sodium for 5 hours, then it is distilled under reduced pressure, to remove alcohols. Isoamyl ether can also be dried with  $\text{CaCl}_2$  and fractionally distilled from  $\text{P}_2\text{O}_5$ . [*Beilstein* **1** IV 1682.]

**Isobutane (2-methylpropane)** [75-28-5] **M 58.1, b -10.2°, d<sub>4</sub><sup>20</sup> 0.557.** Olefins and moisture can be removed by passage at 65° through a bed of silica-alumina catalyst which has previously been evacuated at about 400°. *Alternatively*, water and  $\text{CO}_2$  can be removed by passage through  $\text{P}_2\text{O}_5$ , then asbestos impregnated with  $\text{NaOH}$ . Treatment with anhydrous  $\text{AlBr}_3$  at 0° then removes traces of olefins. Inert gases can be separated by freezing the isobutane at -195° and evacuating out the system. [*Beilstein* **1** IV 282.]

**Isobutene (2-methylpropene, isobutylene)** [115-11-7] **M 56.1, b -6.6°/760mm.** Dry isobutene by passage through anhydrous  $\text{CaSO}_4$  at 0°. Purify it further by freeze-pump-thaw cycles and trap-to-trap distillation. [*Beilstein* **1** IV 796.]

**Isobutyl bromide (1-bromo-2-methylpropane)** [78-77-3] **M 137.0, b 91.2°, d<sub>4</sub><sup>20</sup> 1.260, n<sub>D</sub><sup>20</sup> 1.437.** Partially hydrolyse it to remove any tertiary alkyl halide, then fractionally distil it, then wash it with conc  $\text{H}_2\text{SO}_4$ , water and aqueous  $\text{K}_2\text{CO}_3$ , then redistil it from dry  $\text{K}_2\text{CO}_3$ . [Dunbar & Hammett *J Am Chem Soc* **72** 109 1950, *Beilstein* **1** IV 294.]

**Isobutyl chloride (1-chloro-2-methylpropane)** [513-36-0] **M 92.6, m -131°, 68.8°/760mm, d<sub>4</sub><sup>20</sup> 0.877, n<sub>D</sub><sup>20</sup> 1.398.** Use the same methods as described under *isoamyl chloride*. [*Beilstein* **1** IV 287.]

**Isobutyl chloroformate** [543-27-1] **M 136.6, b 123-127°/atm, 128.8°/atm, d<sub>4</sub><sup>20</sup> 1.053, n<sub>D</sub><sup>20</sup> 1.4070.** It can be dried over  $\text{CaCl}_2$  and fractionated at atmospheric pressure while keeping moisture out. Its purity can be checked by conversion to the *phenyl urethane* derivative with  $\text{PhNCO}$  [Saunders et al. *J Am Chem Soc* **73** 3796 1951.] Its IR (film) has  $\nu_{\text{max}}$  at  $1780\text{cm}^{-1}$ . [Thompson & Jameson *Spectrochim Acta* **13** 236 1959, Röse *Justus Liebigs Ann Chem* **205** 227 1880]. [*Beilstein* **3** IV 26.]

**Isobutyl formate** [542-55-2] **M 102.1, b 98.4°, d<sub>4</sub><sup>20</sup> 0.885, n<sub>D</sub><sup>20</sup> 1.38546.** Wash the formate with saturated aqueous  $\text{NaHCO}_3$ , in the presence of saturated aqueous  $\text{NaCl}$  solution until no further reaction occurs, then with saturated aqueous  $\text{NaCl}$ , dry ( $\text{MgSO}_4$ ) and fractionally distil it. [*Beilstein* **2** H 21, **2** I 18, **2** II 30, **2** III 41, **2** IV 29.]

**Isobutyl iodide (1-iodo-2-methylpropane)** [513-38-2] **M 184.0, b 83°/250mm, 120°/760mm, d<sub>4</sub><sup>20</sup> 1.60, n<sub>D</sub><sup>20</sup> 1.495.** Shake the iodide with conc  $\text{H}_2\text{SO}_4$ , and wash it with water, aqueous  $\text{Na}_2\text{SO}_3$ , and water, dry with  $\text{MgSO}_4$  and distil it. *Alternatively*, pass it through a column of activated alumina before distillation. Store it under nitrogen with mercury in a brown bottle or in the dark. [*Beilstein* **1** IV 299.]

**Isobutyl vinyl ether** [109-53-5] **M 100.2, b 108-110°, d<sub>4</sub><sup>20</sup> 0.768, n<sub>D</sub><sup>20</sup> 1.398.** Wash the ether three times with equal volumes of aqueous 1%  $\text{NaOH}$ , dry it with  $\text{CaH}_2$ , reflux it with sodium for several hours, then fractionally distil it from sodium. [*Beilstein* **1** IV 2054.]

**Isobutyraldehyde** [78-84-2] **M 72.1, b 62.0°, d<sub>4</sub><sup>20</sup> 0.789, n<sub>D</sub><sup>20</sup> 1.377.** Dry isobutyraldehyde with  $\text{CaSO}_4$  and use it immediately after distillation under nitrogen because of the great difficulty in preventing oxidation. It can be purified through its acid bisulfite derivative. [*Beilstein* **1** IV 3262.]

**Isobutyramide (2-methylpropionamide)** [563-83-7] **M 87.1, m 128-129°, b 217-221°/760mm.** Crystallise the amide from acetone, \*benzene,  $\text{CHCl}_3$ ,  $\text{EtOAc}$  or water, then dry it under vacuum over  $\text{P}_2\text{O}_5$  or 99% at  $\text{H}_2\text{SO}_4$ , or at 70°/3 hours in a desiccator. Sublime it under vacuum. [Kent & McAlvain *Org Synth Coll Vol III*

491 1955, *Beilstein* 2 H 293, 2 I 129, 2 II 262, 2 III 654, 2 IV 852.]

**Isobutyric acid (2-methylpropionic acid)** [79-31-2] M 88.1, b 78°/34mm, 154-154.5°/760mm,  $d_4^{20}$  0.949,  $n_D^{20}$  1.393,  $pK^{25}$  4.60. Distil the acid from  $KMnO_4$ , then redistil it from  $P_2O_5$ . [*Beilstein* 2 H 288, 2 I 126, 2 II 257, 2 III 637, 2 IV 843.]

**Isobutyronitrile (2-methylpropionitrile, isopropyl cyanide)** [78-82-0] M 69.1, b 103.6°/760mm,  $d_4^{25}$  0.7650,  $n_D^{20}$  1.378. Shake the nitrile with conc HCl (to remove isonitriles), then with water and aqueous  $NaHCO_3$ . After a preliminary drying with silica gel or Linde type 4A molecular sieves, it is shaken or stirred with  $CaH_2$  until hydrogen evolution ceases, then decanted and distilled from  $P_2O_5$  (not more than 5g/L, to minimise gel formation) or Drierite (b 101-103°/760mm). Finally it is refluxed with, and slowly distilled from  $CaH_2$  (5g/L), taking precautions to exclude moisture. [*Beilstein* 2 H 294, 2 I 129, 2 II 263, 2 III 655, 2 IV 853.]

**Isonitrosoacetone (anti-pyruvic aldehyde-1-oxime)** [31915-82-9] M 87.1, m 65-67°, 69°,  $pK^{25}$  8.3. Crystallise isonitrosoacetone from  $C_6H_6$ , ether/petroleum ether or  $CCl_4$ . It sublimes at 90-100° (water bath temperature)/0.05mm [Kahovec & Kohlrausch *Chem Ber* 75 1547 1942]. It forms an iron and a  $Cu^{2+}$  salt (m 170° dec). [*Beilstein* 1 H 763, 1 I 396, 1 II 822, 1 III 3092, 1 IV 3632.]

**(±)-Isononane (3,3,4-trimethylhexane)** [34464-40-9] M 128.3, b 142°/760mm, 142°/760mm,  $d_4^{20}$  0.7454,  $n_D^{20}$  1.4178. Isononane is passed through columns of activated silica gel and basic alumina (activity 1) and distilled under high vacuum from Na/K alloy. [*Beilstein* 1 III 517, 1 IV 462.]

**Isopentyl formate** [110-45-2] M 116.2, b 27°/10mm, 123-123.6°/760mm, 123-124°/atm,  $d_4^{20}$  0.8713,  $n_D^{20}$  1.391. The colourless liquid ester is soluble in 300 volumes of  $H_2O$  and is soluble in common organic solvents. It is purified by repeated distillation using an efficient column at atmospheric pressure. [*Beilstein* 2 H 22, 2 I 18, 2 II 31, 2 III 43, 2 IV 30.]

**Isoprene (2-methyl-1,3-butadiene)** [78-79-5] M 68.1, b 34.5-35°/762mm,  $d_4^{20}$  0.681,  $n^{25}$  1.4225. Reflux it with sodium then distil it from sodium or  $NaBH_4$  under nitrogen, and pass it through a column containing KOH,  $CaSO_4$  and silica gel. *tert*-Butylcatechol (0.02% w/w) is added, and the isoprene is stored in this way until redistilled before use. The inhibitor (*tert*-butylcatechol) in isoprene can be removed by several washings with dilute NaOH and water. The isoprene is then dried over  $CaH_2$ , distilled under nitrogen at atmospheric pressure, and the fraction distilling at 32° is collected. Store it under nitrogen at -15°. [*Beilstein* 1 H 252, 1 IV 1001.]

**Isopropenyl acetate (1-methylvinyl acetate, 2-acetoxypropene)** [108-22-5] M 100.1, b 96.7°/749mm, 97.2-97.4°/757mm,  $d_4^{25}$  0.909,  $n_D^{20}$  1.4010. Prepared by reacting ketene with a mixture of  $Me_2CO$  and polyphosphoric acid [Wacher US Patent 2 867 653 1957], fractionally distilling through a 75cm packed column, and collecting the fraction with b 96.7°/749mm. It forms an azeotrope with  $H_2O$ . Its equivalent weight is determined by hydrolysis in 90% EtOH containing twice the required amount of NaOH and titrating the excess of alkali; which should give an equivalent weight of ~100.1. Its heat of hydrolysis has been determined [Sunner *Acta Chem Scand* 11 1757 1957]. It is a good source the  $=CMe_2$  group, as in the preparation of Meldrum's acid (see [2033-24-1]), because it is the activated enol ester of acetone. Its FT-IR (neat) has  $\nu_{max}$  at 1757.3, 1673.0, 1431.4, 1372.5, 1253.1, 1198.2, 1027.9, 900.0 and 871.4  $cm^{-1}$ ; the  $^1H$  NMR (60MHz,  $CDCl_3$ , TMS) has  $\delta$  at 1.92 (s, 3H, ester Me), 2.13 (s, 3H, isoPr-Me) and 7.00 (d,  $J = \sim 6Hz$ , 2H,  $=CH_2$ ); and the  $^{13}C$  NMR (15MHz,  $CDCl_3$ , TMS) has  $\delta$  at 19.54, 21.05, 102.05, 152.94 and 169.04. [*Beilstein* 2 III 278, 2 IV 179.]

**Isopropanol (propan-2-ol)** [67-63-0] M 60.1, b 82.5°,  $d_4^{20}$  0.783,  $n^{25.8}$  1.3739,  $pK^{25}$  17.1. Isopropyl alcohol is prepared commercially by dissolution of propene in  $H_2SO_4$ , followed by hydrolysis of the sulfate ester. Major impurities are water, lower alcohols and oxidation products such as aldehydes and ketones. Purification of isopropanol follows substantially the same procedure as for *n*-propyl alcohol. Isopropanol forms a constant-boiling mixture, b 80.3°, with water. Most of the water can be removed from this 91% isopropanol by refluxing with CaO (200g/L) for several hours, then distilling. The distillate can be dried further with  $CaH_2$ , magnesium ribbon, BaO,  $CaSO_4$ , calcium, anhydrous  $CuSO_4$  or Linde type 5A molecular

sieves. Distillation from sulfanilic acid removes ammonia and other basic impurities. Peroxides [indicated by liberation of iodine from weakly acid (HCl) solutions of 2% KI] can be removed by refluxing with solid stannous chloride or with NaBH<sub>4</sub> then the alcohol is fractionally distilled. To obtain isopropanol containing only 0.002M of water, sodium (8g/L) is dissolved in material dried by distillation from CaSO<sub>4</sub>. Isopropyl benzoate (35ml) is then added and, after refluxing for 3 hours, the alcohol is distilled through a 50-cm Vigreux column [Hine & Tanabe *J Am Chem Soc* **80** 3002 1958]. Other purification steps for isopropanol include refluxing with solid aluminium isopropoxide, refluxing with NaBH<sub>4</sub> for 24 hours, and removing acetone by treatment with, and distillation from 2,4-dinitrophenylhydrazine. Peroxides re-form in isopropanol if it is kept for several days in contact with air. [*Beilstein* **1** IV 1461.]

**Isopropyl acetate** [108-21-4] **M 102.1, b 88.4°, d<sub>4</sub><sup>20</sup> 0.873, n<sub>D</sub><sup>20</sup> 1.3773.** Wash the acetate with 50% aqueous K<sub>2</sub>CO<sub>3</sub> (to remove acid), then with saturated aqueous CaCl<sub>2</sub> (to remove any alcohol). Dry it with CaCl<sub>2</sub> and fractionally distil it. [*Beilstein* **2** IV 141.]

**Isopropyl acrylamide** [2210-25-5] **M 113.2, m 60-63°, b 89-92°/2mm, 110-115°/15mm.** Fractionate the amide under reduced pressure, and recrystallise the solid distillate from hexane (**m** 59°), \*C<sub>6</sub>H<sub>6</sub> (**m** 62°) or \*C<sub>6</sub>H<sub>6</sub>/hexane (**m** 62-63°). Store it with 0.05% of 4-*tert*-butylcatechol. It is used for making water soluble swellable hydrogels. [*Beilstein* **4** IV 517.]

**Isopropyl bromide (2-bromopropane)** [75-26-3] **M 123.0, b 0°/69.2mm, 59.4°/760mm, d<sub>4</sub><sup>20</sup> 1.31, n<sub>D</sub><sup>15</sup> 1.42847, n<sub>D</sub><sup>20</sup> 1.4251.** Wash the bromide with 95% H<sub>2</sub>SO<sub>4</sub> (concentrated acid partially oxidised it) until a fresh portion of acid did not become coloured after several hours, then with water, aqueous NaHSO<sub>3</sub>, aqueous 10% Na<sub>2</sub>CO<sub>3</sub> and again with water. (The H<sub>2</sub>SO<sub>4</sub> can be replaced by conc HCl.) Prior to this treatment, isopropyl bromide has been purified by bubbling a stream of oxygen containing 5% ozone through it for 1 hour, followed by shaking with 3% hydrogen peroxide solution, neutralising with aqueous Na<sub>2</sub>CO<sub>3</sub>, washing with distilled water and drying. *Alternatively*, it has been treated with elemental bromine and stored for 4 weeks, then extracted with aqueous NaHSO<sub>3</sub> and dried with MgSO<sub>4</sub>. After the acid treatment, isopropyl bromide can be dried with Na<sub>2</sub>SO<sub>4</sub>, MgSO<sub>4</sub> or CaH<sub>2</sub>, and fractionally distilled. [*Beilstein* **1** IV 208.]

**Isopropyl chloride (2-chloropropane)** [75-29-6] **M 78.5, b 34.8°, d<sub>4</sub><sup>20</sup> 0.864, n<sub>D</sub><sup>20</sup> 1.3779, n<sub>D</sub><sup>25</sup> 1.3754.** Purify the chloride with 95% H<sub>2</sub>SO<sub>4</sub> as described for *isopropyl bromide*, then dry with MgSO<sub>4</sub>, P<sub>2</sub>O<sub>5</sub> or CaH<sub>2</sub>, and fractionally distil it from Na<sub>2</sub>CO<sub>3</sub> or CaH<sub>2</sub>. *Alternatively*, a stream of oxygen containing *ca* 5% ozone is passed through the chloride for about three times as long as is necessary to obtain the first coloration of starch iodide paper by the exit gas, and the liquid is then washed with NaHCO<sub>3</sub> solution to hydrolyse ozonides and remove organic acids before drying and distilling. [*Beilstein* **1** IV 191.]

**Isopropyl ether (diisopropyl ether)** [108-20-3] **M 102.2, b 68.3°, d<sub>4</sub><sup>20</sup> 0.719, n<sub>D</sub><sup>20</sup> 1.3688, n<sub>D</sub><sup>25</sup> 1.36618.** Common impurities are water and peroxides [detected by the liberation of iodine from weakly acid (HCl) solutions of 2% KI]. Peroxides can be removed by shaking with aqueous Na<sub>2</sub>SO<sub>3</sub> or with acidified ferrous sulfate (0.6g FeSO<sub>4</sub> and 6ml conc H<sub>2</sub>SO<sub>4</sub> in 110ml of water, using 5-10g of solution per L of ether), or aqueous NaBH<sub>4</sub> solution. The ether is then washed with water, dried with CaCl<sub>2</sub> and distilled. *Alternatively*, refluxing with LiAlH<sub>4</sub> or CaH<sub>2</sub>, or drying with CaSO<sub>4</sub>, then passage through an activated alumina column, can be used to remove water and peroxides. Other dehydrating agents used with isopropyl ether include P<sub>2</sub>O<sub>5</sub>, sodium amalgam and sodium wire. (The ether is often stored in brown bottles, or in the dark, with sodium wire.) Bonner and Goishi (*J Am Chem Soc* **83** 85 1961) treated isopropyl ether with dilute sodium dichromate/sulfuric acid solution, followed by repeated shaking with a 1:1 mixture of 6M NaOH and saturated KMnO<sub>4</sub>. The ether is washed several times with water, dilute aqueous HCl, and water, with a final washing with, and storage over, ferrous ammonium sulfate acidified with H<sub>2</sub>SO<sub>4</sub>. Blaustein and Gryder (*J Am Chem Soc* **79** 540 1957), after washing with alkaline KMnO<sub>4</sub>, then water, treated the ether with ceric nitrate in nitric acid, and again washed it with water. Hydroquinone is added before drying with CaCl<sub>2</sub> and MgSO<sub>4</sub>, and refluxing with sodium amalgam (108g Hg/100g Na) for 2 hours under nitrogen. The distillate (nitrogen atmosphere) is made 2 x 10<sup>-5</sup>M in hydroquinone to inhibit formation of peroxides (which is negligible if the ether is stored in the dark). Catechol (pyrocatechol) and resorcinol are alternative inhibitors. [*Beilstein* **1** IV 1471.]

**Isopropyl iodide (2-iodopropane)** [75-30-9] **M 170.0, b 88.9°, d<sub>4</sub><sup>20</sup> 1.70, n<sub>D</sub><sup>20</sup> 1.4987.** Treat the iodide with bromine, followed by extraction of free halogen with aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> or NaHSO<sub>3</sub>, washing with water, drying (MgSO<sub>4</sub> or CaCl<sub>2</sub>) and distilling. (The treatment with bromine is optional.) Other purification methods include passage through activated alumina, or shaking with copper powder or mercury to remove iodine, drying with P<sub>2</sub>O<sub>5</sub> and distilling. Washing with conc H<sub>2</sub>SO<sub>4</sub> or conc HCl (to remove any alcohol), water, aqueous Na<sub>2</sub>SO<sub>3</sub>, water and aqueous Na<sub>2</sub>CO<sub>3</sub> has also been used. Treatment with silica gel causes some liberation of iodine. Distillations should be carried out at slightly reduced pressure. Purified isopropyl iodide is stored in the dark in the presence of a little mercury. [*Beilstein* 1 IV 223.]

**Isopropyl methyl ether** [598-53-8] **M 74.1, b 32.5°/777mm, d<sup>15</sup> 0.724, n<sub>D</sub><sup>20</sup> 1.3576.** Purify the ether by drying with CaSO<sub>4</sub>, passing through a column of alumina (to remove peroxides) and fractional distillation. [*Beilstein* 1 H 362, 1 II 381, 1 III 1458, 1 IV 1471.]

**Isovaleric acid (3-methylbutyric acid)** [502-74-2] **M 102.1, b 176.5°/762mm, d<sub>4</sub><sup>20</sup> 0.927, n<sup>15</sup> 1.4064, n<sub>D</sub><sup>20</sup> 1.40331, pK<sup>25</sup> 4.77.** Dry the acid (Na<sub>2</sub>SO<sub>4</sub>), then fractionally distil. [*Beilstein* 2 IV 895.]

**Itaconic acid (2-propen-1,2-dicarboxylic acid)** [97-65-4] **M 130.1, m 165-166°, pK<sub>1</sub><sup>25</sup> 3.63, pK<sub>2</sub><sup>25</sup> 5.00.** Crystallise itaconic acid from EtOH, EtOH/water or EtOH/\*benzene. [*Beilstein* 2 IV 2228.]

**Itaconic anhydride (2-propen-1,2-dicarboxylic anhydride)** [2170-03-8] **M 112.1, m 66-68°, 67-68°, 68°, b 139-140°/30mm.** Crystallise the anhydride from CHCl<sub>3</sub>/petroleum ether. It can be distilled under reduced pressure. Distillation at atmospheric pressure, or prolonged distillation causes rearrangement to citraconic anhydride (2-methylmaleic anhydride). If the material (as seen in the IR spectrum) contains much free acid, then heat with acetyl chloride or SOCl<sub>2</sub>, evaporate and distil at as high a vacuum as possible. The crude anhydride deposits crystals of itaconic acid on standing probably due to hydrolysis by H<sub>2</sub>O — store it in sealed ampoules under dry N<sub>2</sub>. [Skinner et al. *Org Synth Coll Vol II* 368 1943, IR: Nagai *Bull Chem Soc Jpn* 37 369 1964, Kelly & Segura *J Am Chem Soc* 56 2497 1934, *Beilstein* 17/11 V 66.]

**Kerosene** [8008-20-6] (mixture of hydrocarbons) **b ~175-225°, ~190-250°, d<sub>4</sub><sup>20</sup> 0.75-0.82, n<sub>D</sub><sup>20</sup> 1.443.** Stir it with conc H<sub>2</sub>SO<sub>4</sub> until a fresh portion of acid remains colourless, then wash with water, dry with solid KOH and distil it in a Claisen flask. For more complete drying, the kerosene can be refluxed with Na, and distilled from Na.

**Ketene** [463-51-4] **M 42.0, b -56°, -41°, d<sub>4</sub><sup>20</sup> 1.093, n<sub>D</sub><sup>20</sup> 1.441.** Ketene is prepared by pyrolysis of acetic anhydride. Purify it by passing through a trap at -75° and collecting in a liquid-nitrogen-cooled trap. Ethylene is removed by evacuating the ethylene in an isopentane-liquid-nitrogen slush pack at -160°. Store it at room temperature in a suitable container in the dark or better at -80°, but do not store it under pressure as it may **EXPLODE**. It is a strong **IRRITANT** when inhaled and is as **POISONOUS** as phosgene. See diketene in "Heterocyclic Compounds", in this Chapter. [Hurd *Org Synth Coll Vol I* 330 1941, Andreades & Carlson *Org Synth Coll Vol V* 679 1973.]

**L(+)-Lactic acid (S(+)-2-hydroxypropionic acid)** [79-33-4] **M 90.1, m 52.8°, b 105°/0.1mm, [α]<sub>D</sub><sup>20</sup> +3.82° (H<sub>2</sub>O), pK<sup>31</sup> 3.83.** Purify lactic acid by fractional distillation at 0.1mm pressure, followed by fractional crystallisation from diethyl ether/isopropyl ether (1:1, dried with sodium). [Borsook et al. *J Biol Chem* 102 449 1933.] The solvent mixture, \*benzene/diethyl ether (1:1) containing 5% petroleum ether (b 60-80°) has also been used. [Brin *Biochemical Preparations* 3 61 1953, *Beilstein* 3 IV 633.]

**Lanthanide shift reagents** A variety of these reagents are available commercially, and they are generally quite stable and should not deteriorate on long storage in a dry state and in the absence of light. [See J.R. Campbell *Aldrichimica Acta* 4 55 1971, G.R. Sullivan in *Top Stereochem* (Eliel & Allinger Eds) J Wiley & Sons Vol 10 287 1978, T.C. Morrill Ed. *Lanthanide Shift Reagents* Deerfield Beach Florida 1986, ISBN 0895731193.]

**Lauraldehyde (1-dodecanal)** [112-54-9] **M 184.3, b 99.5-100°/3.5mm, n<sub>D</sub><sup>24.7</sup> 1.4328.** Convert lauraldehyde to the bisulfite addition compound by shaking with saturated aqueous NaHSO<sub>3</sub> for 1 hour. The precipitate is filtered off, washed with ice cold water, EtOH and ether, then decomposed with aqueous Na<sub>2</sub>CO<sub>3</sub>. The aldehyde is extracted into diethyl ether which, after drying and evaporating, gives an oil which is fractionally distilled under vacuum. [Beilstein 1 IV 3380.]

**Lauric acid (1-dodecanoic acid)** [143-07-7] **M 200.3, m 44.1°, b 141-142°/0.6-0.7mm, 225°/100mm, pK<sub>a</sub><sup>20</sup> 5.3.** Distil the acid in a vacuum. Also crystallise it from absolute EtOH, or from acetone at -25°. *Alternatively,* purify it *via* its *methyl ester* (b 140.0°/15mm), as described for *capric acid* (see [334-48-5]). It has also been purified by zone melting. [cf Beilstein 1 III 2913.]

**Lauryl peroxide (di-dodecyl peroxide)** [105-74-8] **M 398.6, m 53-54°.** Crystallise it from *n*-hexane or \*benzene and store it below 0°. Potentially **EXPLOSIVE**. [cf Beilstein 2 IV 1102.]

**Z-Maleamic acid (cis-maleic acid monoamide)** [557-24-4] **M 115.1, m 172°, 172-173°(dec), 178-180°, pK<sub>Est</sub> ~2.65.** Crystallise it from EtOH. [Beilstein 2 H 752, 2 II 646, 2 III 1927, 2 IV 1927.] **IRRITANT.**

**Maleic acid** [110-16-7] **M 116.1, m 143.5°, pK<sub>1</sub><sup>25</sup> 1.91, pK<sub>2</sub><sup>25</sup> 6.33.** Crystallise the acid from acetone/petroleum ether (b 60-80°) or hot water. Dry it at 100°. [Beilstein 2 H 748, 2 I 303, 2 II 641, 2 III 1911, 2 IV 2199.]

**Maleic anhydride** See furan-2,5-dione in "Heterocyclic Compounds", in this Chapter.

**Maleic hydrazide** [123-33-1] **M 112.1, m 144°(dec), pK<sub>1</sub><sup>25</sup> 5.67, pK<sub>2</sub><sup>25</sup> 13.3.** Crystallise the hydrazide from water. Dry it at ~100° over P<sub>2</sub>O<sub>5</sub>. [Beilstein 24 III/IV 1186.]

**Maleimide** See pyrrol-2,5-dione in "Heterocyclic Compounds", in this Chapter.

**Maleuric acid (Z-N-carbamoylmaleamic acid)** [105-61-3] **M 158.1, m 167-168°(dec).** Crystallise the acid from hot water. Dry it at ~100° over H<sub>2</sub>SO<sub>4</sub>. [Batt et al. *J Am Chem Soc* 76 3663 1954.]

**dl-Malic acid** [617-48-1 and 6915-15-7] **M 134.1, m 128-129°.** Crystallise the acid from acetone, then from acetone/CCl<sub>4</sub>, or from ethyl acetate by adding petroleum ether (b 60-70°). Dry it at 35° under 1mm pressure to avoid formation of the anhydride. [Beilstein 3 IV 1124.]

**L-Malic acid (S(-)-2-hydroxysuccinic acid)** [97-67-6] **M 134.1, m 104.5-106°, [α]<sub>D</sub><sup>20</sup> -2.3° (c 8.5, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> -30° (c 5.5, pyridine), pK<sub>1</sub><sup>25</sup> 3.46, pK<sub>2</sub><sup>25</sup> 5.10.** Crystallise *S*-malic acid (charcoal) from ethyl acetate/petroleum ether (b 55-56°), keeping the temperature below 65°. Or dissolve it by refluxing in fifteen parts of anhydrous diethyl ether, decant, concentrate to one-third volume and crystallise it at 0°, repeatedly to constant melting point. [Beilstein 3 IV 1123.]

**Malonamide** [108-13-4] **M 102.1, m 170°.** Crystallise the amide from water. [Beilstein 2 IV 1887.]

**Malonic acid** [141-82-2] **M 104.1, m 136°, pK<sub>1</sub><sup>25</sup> 2.58, pK<sub>2</sub><sup>25</sup> 5.69.** Crystallise malonic acid from \*benzene/diethyl ether (1:1) containing 5% of petroleum ether (b 60-80°), wash with diethyl ether, then recrystallise it from H<sub>2</sub>O or acetone. Dry it under vacuum over conc H<sub>2</sub>SO<sub>4</sub>. [Beilstein 2 IV 1874.]

**Malonitrile** [109-77-3] **M 66.1, m 32-34°, b 109°/20mm, 113-118°/25mm, 220°/760mm.** Crystallise the nitrile from water, EtOH, \*benzene or chloroform. Distil it in a vacuum from, and store over, P<sub>2</sub>O<sub>5</sub>. [Bernasconi et al. *J Am Chem Soc* 107 7692 1985, Gratenhuis *J Am Chem Soc* 109 8044 1987, Beilstein 2 IV 1892.]

**Meprobamate [2,2-di(carbamoyloxymethyl)pentane]** [57-53-4] **M 246.3, m 104-106°**. Crystallise it from hot water, aqueous EtOH (**m 104-105.5°**) or xylene (**m 104.1-105.3°**). It can be an addictive drug. [*Beilstein 3 IV 73.*]

**2-Mercaptoethanol** [60-24-2] **M 78.1, b 44°/4mm, 53.5°/10mm, 58°/12mm, 68°/20mm, 78.5°/40mm, 96-97° (92°)/100mm, 157°/748mm, d<sub>4</sub><sup>20</sup> 1.114, n<sub>D</sub><sup>20</sup> 1.500, pK<sup>25</sup> 9.72 (9.43)**. Purify it by distilling in a vacuum. Distilling at atmospheric pressure causes some oxidation and should be done in an inert atmosphere. [Woodward *J Chem Soc* 1892 1948.] **It has a foul odour, is irritating to the eyes, nose and skin — should be handled in an efficient fume cupboard.** It is miscible with H<sub>2</sub>O, EtOH, Et<sub>2</sub>O and \*C<sub>6</sub>H<sub>6</sub> and the UV has λ<sub>max</sub> at 235nm. The 2,4-dinitrophenyl thioether has **m 101-102°** (from EtOH or aqueous MeOH) [Grogen et al. *J Org Chem* **20** 50 1955]. [*Beilstein 1 IV 2428.*]

**Mesaconic acid (methylfumaric acid)** [498-24-8] **M 130.1, m 204-205°, pK<sup>18</sup> 4.82**. Crystallise it from H<sub>2</sub>O or EtOH [Katakis et al. *J Chem Soc, Dalton Trans* 1491 1986]. [*Beilstein 2 IV 2231.*]

**Mesityl oxide (4-methyl-3-penten-2-one)** [141-79-7] **M 98.2, b 57°/55mm, 128-129°/745mm, 112°/760mm, n<sub>D</sub><sup>24</sup> 1.4412, d 0.854, pK<sup>20</sup> -5.36 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>)**. Purify it by distillation, preferably in a vacuum or via the semicarbazone (**m 165°**) which is decomposed to pure ketone. The 2,4-dinitrophenylhydrazone (**m 205-206°**) crystallises from EtOH. [Johnson *J Am Chem Soc* **73** 5888 1951, Johnson *J Am Chem Soc* **75** 2720 1953, Erskine & Waight *J Chem Soc* 3425 1960, *Beilstein 1 H 736, 1 I 382, 1 II 793, 1 III 2995, 1 IV 3471.*]

**α-Methacraldehyde (methacrolein)** [78-85-3] **M 68.1, b 68.4°, d<sub>4</sub><sup>20</sup> 0.849, n<sub>D</sub><sup>20</sup> 1.416**. Fractionally distil it under nitrogen through a short Vigreux column. Store it in sealed ampoules. (Slight polymerisation may occur.) [*Beilstein 1 IV 3455.*]

**Methacrylamide** [79-39-0] **M 85.1, m 111-112°**. Crystallise the amide from \*benzene or ethyl acetate and dry it under vacuum at room temperature. [*Beilstein 2 IV 1538.*]

**Methacrylic acid** [79-41-4] **M 86.1, b 72°/14mm, 160°/760mm, d<sub>4</sub><sup>20</sup> 1.015, n<sub>D</sub><sup>20</sup> 1.431, pK<sup>25</sup> 4.65**. Aqueous methacrylic acid (90%) is saturated with NaCl (to remove the bulk of the water), then the organic phase is dried with CaCl<sub>2</sub> and distilled under vacuum. Polymerisation inhibitors should be added to the distillate and include 0.25% *p*-methoxyphenol, 0.1% hydroquinone, or 0.05% *N,N'*-diphenyl-*p*-phenylenediamine. [*Beilstein 2 IV 1518.*]

**Methacrylic anhydride** [760-93-0] **M 154.2, b 65°/2mm, d<sub>4</sub><sup>20</sup> 1.040, n<sub>D</sub><sup>20</sup> 1.454**. Distil the anhydride at 2mm pressure, immediately before use, in the presence of hydroquinone. [*Beilstein 2 IV 1537.*]

**Methacrylonitrile** [126-98-7] **M 67.1, b 90.3°, d<sub>4</sub><sup>20</sup> 0.800, n<sub>D</sub><sup>20</sup> 1.4007, n<sub>D</sub><sup>30</sup> 1.3954**. Wash it with saturated aqueous NaHSO<sub>3</sub> (to remove inhibitors such as *p-tert*-butylcatechol), 1% NaOH in saturated NaCl and then with saturated NaCl. Dry it with CaCl<sub>2</sub> and fractionally distil it under nitrogen to separate it from impurities such as methacrolein and acetone. [*Beilstein 2 IV 1539.*]

**Methacryloyl chloride** [920-46-7] **M 104.5, m -60°, b 95-96°/760mm, 98.4°/772mm, d<sub>4</sub><sup>25</sup> 1.076, n<sub>D</sub><sup>20</sup> 1.4432**. Purify the ester by fractional distillation. If it contains the acid (OH bands in the IR) then add redistilled SOCl<sub>2</sub> (with cooling) and cuprous chloride (ca to 2%), reflux the mixture gently for 1 hour and fractionate it through a 1 metre column packed with glass helices. Redistillation then provides the acid chloride in high purity as a colourless liquid. It is necessary to keep the apparatus moisture free (use CaCl<sub>2</sub> tubes). Stabilise it with 0.05% of 2,6-di-*tert*-butyl-4-methylphenol. [Lal & Green *J Org Chem* **20** 1032 1955, *Beilstein 2 IV 1537.*]

**Methane** [74-82-8] **M 16.0, m -184°, b -164°/760mm, -130°/6.7atmospheres, d<sub>4</sub><sup>0</sup> 0.554 (cf d<sub>4</sub><sup>0</sup> 1.00 for air)**. Dry methane by passing over CaCl<sub>2</sub> and P<sub>2</sub>O<sub>5</sub>, then through a Dry-ice trap and fractionally distil it from a liquid-nitrogen trap. Oxygen can be removed by prior passage in a stream of hydrogen over reduced copper oxide at 500°, and higher hydrocarbons can be removed by chlorinating about 10% of the sample: the hydrocarbons, chlorides and HCl are readily separated from the methane by condensing the sample in the liquid nitrogen trap

and fractionally distilling it. Methane has also been washed with conc H<sub>2</sub>SO<sub>4</sub>, then solid NaOH and then 30% NaOH solution. It is dried with CaCl<sub>2</sub>, then P<sub>2</sub>O<sub>5</sub>, and condensed in a trap at liquid air temperature, then transferred to another trap cooled in liquid nitrogen. CO<sub>2</sub>, O<sub>2</sub>, N<sub>2</sub> and higher hydrocarbons can be removed from methane by adsorption on charcoal. [Eiseman & Potter *J Res Nat Bur Stand* **58** 213 1957, *Beilstein* **1** IV 3.] **HIGHLY FLAMMABLE.**

**Methanesulfonic acid** [75-75-2] **M 96.1, m 20°, b 134.5-135°/3mm, d<sub>4</sub><sup>20</sup> 1.483, n<sub>D</sub><sup>20</sup> 1.432, pK<sup>25</sup> -1.86 (-1.2).** Dry the acid, either by azeotropic removal of water with \*benzene or toluene, or by stirring 20g of P<sub>2</sub>O<sub>5</sub> with 500ml of the acid at 100° for 0.5 hours. Then distil it under vacuum and fractionally crystallise it by partial freezing. Sulfuric acid, if present, can be removed by prior addition of Ba(OH)<sub>2</sub> to a dilute solution, filtering off the BaSO<sub>4</sub> and concentrating under reduced pressure; and is sufficiently pure for most applications. [*Beilstein* **4** IV 10.]

**Methanesulfonothioic acid Na salt (sodium methanethiosulfonate, sodium methylthiosulfonate)** [1950-85-2] **M 134.1, m 265°(dec).** Recrystallise the salt from H<sub>2</sub>O (plates as *monohydrate*) or MeOH. The *potassium salt* crystallises from H<sub>2</sub>O, EtOH or MeOH (thick plates) with **m 201-202°** [Foss *Acta Chem Scand* **10** 868 1956]. The *S-benzylisothiuronium salt* has **m 141-142°** (from EtOH) [Kurzer & Powell *J Chem Soc* 3733 1952]. [*Beilstein* **4** IV 31.] It is used for preparing unsymmetrical disulfides [Grayson et al. *J Org Chem* **70** 9740 2005, Kalai et al. *Synthesis* 439 2006].

**Methanesulfonyl chloride** [124-63-0] **M 114.5, b 55°/11mmm, d<sub>4</sub><sup>20</sup> 1.474, n<sub>D</sub><sup>20</sup> 1.452.** Distil the sulfonyl chloride from P<sub>2</sub>O<sub>5</sub> under vacuum. It is a strong **IRRITANT**. [*Beilstein* **4** IV 27.]

**Methanol** [67-56-1] **M 32.0, b 64.5°, d<sup>15</sup> 0.79609, d<sup>25</sup> 1.32663, n<sup>15</sup> 1.33057, n<sup>25</sup> 1.32663, pK<sup>25</sup> 15.5.** Almost all methanol is now obtained synthetically. Likely impurities are water, acetone, formaldehyde, ethanol, methyl formate and traces of dimethyl ether, methylal, methyl acetate, acetaldehyde, carbon dioxide and ammonia. Most of the water (down to about 0.01%) can be removed by fractional distillation. Drying with CaO is unnecessary and wasteful. Anhydrous methanol can be obtained from "absolute" material by passage through Linde type 4A molecular sieves, or by drying with CaH<sub>2</sub>, CaSO<sub>4</sub>, or with just a little more sodium than required to react with the water present, in all cases the methanol is then distilled. Two treatments with sodium reduces the water content to about 5 x 10<sup>-5</sup>%. [Friedman et al. *J Am Chem Soc* **83** 4050 1961.] Lund and Bjerrum [*Chem Ber* **64** 210 1931] warmed clean dry magnesium turnings (5g) and iodine (0.5g) with 50-75ml of "absolute" methanol in a flask until the iodine disappeared and all the magnesium was converted to the methoxide. Up to 1L of methanol was added and, after refluxing for 2-3 hours, it was distilled off, excluding moisture from the system. Redistillation from tribromobenzoic acid removes basic impurities and traces of magnesium oxides, and leaves conductivity-quality material. The method of Hartley and Raikes [*J Chem Soc* **127** 524 1925] gives a slightly better product. This consists of an initial fractional distillation, followed by distillation from aluminium methoxide, and then ammonia and other volatile impurities are removed by refluxing for 6 hours with freshly dehydrated CuSO<sub>4</sub> (2g/L) while dry air is passed through: the methanol is finally distilled. (The aluminium methoxide is prepared by warming with aluminium amalgam (3g/L) until all the aluminium has reacted. The amalgam is obtained by warming pieces of sheet aluminium with a solution of HgCl<sub>2</sub> in dry methanol.) This treatment also removes aldehydes.

If acetone is present in the methanol, it is usually removed prior to drying. Bates, Mullaly and Hartley [*J Chem Soc* 401 1923] dissolved 25g of iodine in 1L of methanol and then poured the solution, with constant stirring, into 500ml of M NaOH. Addition of 150ml of water precipitated iodoform. The solution was allowed to stand overnight, filtered, then boiled under reflux until the odour of iodoform disappeared, and fractionally distilled: (This treatment also removes formaldehyde.) Morton and Mark [*Ind Eng Chem (Anal Edn)* **6** 151 1934] refluxed methanol (1L) with furfural (50ml) and 10% NaOH solution (120ml) for 6-12 hours, the refluxing resin carries down with it the acetone and other carbonyl-containing impurities. The alcohol was then fractionally distilled. Evers and Knox [*J Am Chem Soc* **73** 1739 1951], after refluxing 4.5L of methanol for 24 hours with 50g of magnesium, distilled off 4L of it, which they then refluxed with AgNO<sub>3</sub> for 24 hours in the absence of moisture or CO<sub>2</sub>. The methanol was again distilled, shaken for 24 hours with activated alumina before being filtered through a glass sinter and distilled under nitrogen in an all-glass still. Material suitable for conductivity

work was obtained.

Variations of the above methods have also been used. For example, a sodium hydroxide solution containing iodine has been added to methanol and, after standing for 1 day, the solution has been poured slowly into about a quarter of its volume of 10% AgNO<sub>3</sub>, shaken for several hours, then distilled. Sulfanilic acid has been used instead of tribromobenzoic acid in Lund and Bjerrum's method. A solution of 15g of magnesium in 500ml of methanol has been heated under reflux, under nitrogen, with hydroquinone (30g), before degassing and distilling the methanol, which was subsequently stored with magnesium (2g) and hydroquinone (4g per 100ml). Refluxing for about 12 hours removes the bulk of the formaldehyde from methanol: further purification has been obtained by subsequent distillation, refluxing for 12 hours with dinitrophenylhydrazine (5g) and H<sub>2</sub>SO<sub>4</sub> (2g/L), and again fractionally distilling. [*Beilstein* 1 IV 1227.]

**Rapid purification:** Methanol purification is the same as for ethanol. Another simple purification procedure consists of adding 2g of NaBH<sub>4</sub> to 1.5L methanol, gently bubbling argon through it and refluxing for a day at 30°, then adding 2g of freshly cut sodium (washed with methanol) and refluxing for 1 day before distilling. The middle fraction is taken. [Jou & Freeman *J Phys Chem* 81 909 1977.]

**Methoxyacetic acid** [625-45-6] **M 90.1, b 97°/13-14mm, d<sub>4</sub><sup>20</sup> 1.175, n<sub>D</sub><sup>20</sup> 1.417, pK<sup>25</sup> 3.57.** Fractionally crystallise the acid by repeated partial freezing, then fractionally distil it under vacuum through a vacuum-jacketed Vigreux column ~20cm long. [*Beilstein* 3 IV 574.]

**Methoxyamine hydrochloride** [593-56-6] **M 83.5, m 151-152°, pK<sup>25</sup> 4.60.** Crystallise the hydrochloride from absolute EtOH or EtOH by addition of diethyl ether. [Kovach et al. *J Am Chem Soc* 107 7360 1985, *Beilstein* 1 IV 1252.]

**2-Methoxyethanol (methylcellosolve)** [109-86-4] **M 76.1, b 124.4°, d<sub>4</sub><sup>20</sup> 0.964, n<sub>D</sub><sup>20</sup> 1.4017, pK<sup>25</sup> 14.8.** Peroxides can be removed by refluxing with stannous chloride or by filtration under slight pressure through a column of activated alumina. 2-Methoxyethanol can be dried with K<sub>2</sub>CO<sub>3</sub>, CaSO<sub>4</sub>, MgSO<sub>4</sub> or silica gel, then distilled from sodium. Aliphatic ketones (and water) can be removed by making the solvent 0.1% in 2,4-dinitrophenylhydrazine and allowing to stand overnight with silica gel before fractionally distilling. [*Beilstein* 1 IV 2375.]

**2-Methoxyethoxymethylchloride (MEMCl)** [3970-21-6] **M 124.6, b 50-52°/13mm, 140-145°(dec)/atm, d<sub>4</sub><sup>20</sup> 1.092, n<sub>D</sub><sup>20</sup> 1.427.** Possible impurities are methoxyethanol (b 124°/atm), HCHO and HCl which can be removed below the boiling point of MEMCl. Purify MEMCl by fractional distillation in a vacuum. If too impure, prepare it from methoxyethanol (152g) and *s*-trioxane (66g) by bubbling a stream of dry HCl (with stirring) until a clear mixture is obtained. Dilute with pentane (900ml), dry (3 hours over 100g MgSO<sub>4</sub>, at 5°), evaporate and the residue is distilled in a vacuum. It is MOISTURE SENSITIVE and TOXIC. The MEM.NEt<sub>3</sub><sup>+</sup>Cl<sup>-</sup> salt, prepared by reaction with 1.3 equivalents of Et<sub>3</sub>N (16 hours/25°) and dried in a vacuum, has m 58-61°, and is moisture sensitive. [Corey et al. *Tetrahedron Lett* 809 1976, Yoshimatsu et al. *J Org Chem* 59 1011 1994, Greene & Wuts *Protective Groups in Organic Synthesis* 3<sup>rd</sup> edn, J Wiley & Sons NY 1991.] **CARCINOGEN.**

**2-Methoxyethylamine** [109-85-3] **M 75.1, b 94°, d<sub>4</sub><sup>20</sup> 0.874, n<sub>D</sub><sup>20</sup> 1.407, pK<sup>25</sup> 9.40.** An aqueous 70% solution of the amine is dehydrated by azeotropic distillation with \*benzene or methylene chloride and the amine is distilled twice from zinc dust. Store it in a tight container as it absorbs CO<sub>2</sub> from the atmosphere. [*Beilstein* 4 IV 1411.]

**8-Methoxypsoralen** See xanthotoxin in “Miscellaneous Compounds”, Chapter 7.

***N*-Methylacetamide** [79-16-3] **M 73.1, m 30°, b 70-71°/2.5-3mm, pK<sub>1</sub><sup>25</sup> -3.70, pK<sub>2</sub><sup>25</sup> -0.42.** Fractionally distil it under vacuum, then fractionally crystallise it twice from its melt. Likely impurities include acetic acid, methyl amine and H<sub>2</sub>O. For a detailed purification procedure, see Knecht and Kolthoff, *Inorg Chem* 1 195 1962. Although *N*-methylacetamide is commercially available it is often extensively contaminated with acetic acid, methylamine, water and an unidentified impurity. The recommended procedure is to synthesise it in the laboratory by direct reaction. The gaseous amine is passed into hot glacial acetic acid, to give a partially

aqueous solution of methylammonium acetate which is heated to *ca* 130° to expel water. Chemical methods of purification such as extraction by petroleum ether, treatment with H<sub>2</sub>SO<sub>4</sub>, K<sub>2</sub>CO<sub>3</sub> or CaO can be used but are more laborious.

Tests for purity include the Karl Fischer titration for water; this can be applied directly. Acetic acid and methylamine can be detected polarographically.

In addition to the above, purification of *N*-methylacetamide can be achieved by fractional freezing, including zone melting, repeated many times, or by vacuum distillation under reduced pressures. For details of zone melting techniques, see Knecht in *Recommended Methods for Purification of Solvents and Tests for Impurities*, Coetzee Ed. Pergamon Press 1982. [*Beilstein* 4 IV 176.]

**Methyl acetate** [79-20-9] **M 74.1, b 56.7-57.2°**,  $d_4^{20}$  0.934,  $n_D^{20}$  1.36193,  $n_D^{25}$  1.3538,  $pK^{20}$  -7.28 (**H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>**). Methanol in methyl acetate can be detected by measuring its solubility in water. At 20°, the solubility of methyl acetate in water is *ca* 35g per 100ml, but 1% MeOH confers complete miscibility. Methanol can be removed by conversion to methyl acetate, by refluxing for 6 hours with acetic anhydride (85ml/L), followed by fractional distillation. Acidic impurities can be removed by shaking with anhydrous K<sub>2</sub>CO<sub>3</sub> and distilling. An alternative treatment is with acetyl chloride, followed by washing with concentrated NaCl and drying with CaO or MgSO<sub>4</sub>. (Solid CaCl<sub>2</sub> cannot be used because it forms a crystalline addition compound.) Distillation from copper stearate destroys peroxides. Free alcohol or acid can be eliminated from methyl acetate by shaking with strong aqueous Na<sub>2</sub>CO<sub>3</sub> or K<sub>2</sub>CO<sub>3</sub> (three times), then with aqueous 50% CaCl<sub>2</sub> (three times), saturated aqueous NaCl (twice), drying with K<sub>2</sub>CO<sub>3</sub> and distilling it from P<sub>2</sub>O<sub>5</sub>. [*Beilstein* 2 IV 122.]

**Methyl acetimidate hydrochloride** [14777-27-6] **M 109.6, m 93-95°, 105°(dec)**,  $pK_{Est}$  ~5.5. Crystallise the imidate from methanol by adding dry ether to a ratio of 1:1 and cooling at 0°. Filter off the crystals in a cold room, wash them with methanol/ether (1:2), then dry in a vacuum. [Hunter & Ludwig *J Am Chem Soc* 84 3491 1962.] The *free base* has **b** 90-91°/765mm,  $d_4^{20}$  0.867,  $n_D^{20}$  1.403. [Hunter & Ludwig *Methods Enzymol* 25 585 1973, *Beilstein* 2 IV 181.]

**Methyl acrylate** [96-33-3] **M 86.1, b 80°,  $d_4^{20}$  0.9535,  $n_D^{20}$  1.4040**. Wash the ester repeatedly with aqueous NaOH until free from inhibitors (such as hydroquinone), then wash it with distilled water, dry (CaCl<sub>2</sub>) and fractionally distil it under reduced pressure in an all-glass apparatus. Seal it under nitrogen and store it at 0° in the dark. [Bamford & Han *J Chem Soc, Faraday Trans 1* 78 855 1982, *Beilstein* 2 IV 1457.]

**Methylamine (gas)** [74-89-5] **M 31.1, b -7.55°/719mm,  $pK^{25}$  10.62**. Dry the amine with sodium or BaO. It is commercially available in metal cylinders. [*Beilstein* 4 IV 118.]

**Methylamine hydrochloride** [593-51-1] **M 67.5, m 231.8-233.4°, b 225-230°/15mm,  $pK^{25}$  10.62**. Crystallise the salt from *n*-butanol, absolute EtOH or MeOH/CHCl<sub>3</sub>. Wash it with CHCl<sub>3</sub> to remove traces of dimethylamine hydrochloride. Dry it under vacuum first with H<sub>2</sub>SO<sub>4</sub> then P<sub>2</sub>O<sub>5</sub>. It is deliquescent; store it in a desiccator over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* 4 IV 122.]

**Methyl bromide** [74-83-9] **M 94.9, b 3.6°**. Purify it by bubbling through conc H<sub>2</sub>SO<sub>4</sub>, followed by passage through a tube containing glass beads coated with P<sub>2</sub>O<sub>5</sub>. Also purify it by distillation from AlBr<sub>3</sub> at -80°, by passage through a tower of KOH pellets and by partial condensation. [*Beilstein* 1 IV 68.]

**2-Methylbutane (isopentane)** [78-78-4] **M 72.2, b 27.9°,  $d_4^{20}$  0.621,  $n_D^{20}$  1.35373,  $n_D^{25}$  1.35088**. Stir isopentane for several hours in the cold with conc H<sub>2</sub>SO<sub>4</sub> (to remove olefinic impurities), then wash it with H<sub>2</sub>O, aqueous Na<sub>2</sub>CO<sub>3</sub> and H<sub>2</sub>O again. Dry it with MgSO<sub>4</sub> and fractionally distil it using a Todd column packed with glass helices. Material transparent down to 180nm is obtained by distilling from sodium wire, and passing through a column of silica gel which had previously been dried in place at 350° for 12 hours before use. [Potts *J Phys Chem* 20 809 1952, *Beilstein* 1 IV 320.]

**2-Methyl-1-butanol** [137-32-6, *RS*(±) 34713-94-5, *S*(-) 1565-80-6] **M 88.2, b 130°(*RS*), 128.6°(*S*)**,  $[\alpha]_D^{25}$  -5.8° (neat),  $d_4^{20}$  0.809,  $n_D^{25}$  1.4082. Reflux the butanol with CaO, distil, reflux with magnesium and again

fractionally distil it. A small sample of highly purified material is obtained by fractional crystallisation after conversion into a suitable ester such as the trinitrophthalate or the 3-nitrophthalate. The latter is converted to the cinchonine salt in acetone and recrystallised from  $\text{CHCl}_3$  by adding pentane. The salt is saponified, extracted with ether, and fractionally distilled. [Terry et al. *J Chem Eng Data* **5** 403 1960, *Beilstein* **1** IV 1666.]

**3-Methyl-2-butanol** [598-75-4] **M 88.2, b 111.5°, d<sub>4</sub><sup>20</sup> 0.807, n<sub>D</sub><sup>20</sup> 1.4095, n<sub>D</sub><sup>25</sup> 1.4076.** Reflux it with magnesium, then fractionally distil it. [*Beilstein* **1** IV 1675.]

**3-Methyl-2-butanone (methyl isopropyl ketone)** [563-80-4] **M 86.1, b 93-94°/752mm, d<sub>4</sub><sup>20</sup> 0.818, n 1.410, pK<sup>25</sup> -7.1 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Reflux the ketone with a little  $\text{KMnO}_4$ . Fractionate it through a spinning-band column, dry with  $\text{CaSO}_4$  and distil it. [*Beilstein* **1** IV 3287.]

**2-Methyl-2-butene** see **amylene** above.

**2-Methyl-3-butyne-2-amine (1,1-dimethylpropargylamine, 3-amino-3-methyl-1-butyne)** [2978-58-7] **M 83.1, b 79-80°/760mm, d<sub>4</sub><sup>25</sup> 0.790, n<sub>D</sub><sup>25</sup> 1.4183, pK<sub>Est</sub> ~8.0.** Dissolve the amine in  $\text{Et}_2\text{O}$ , dry over anhydrous  $\text{K}_2\text{CO}_3$ , filter, evaporate and distil (preferably under  $\text{N}_2$ ). Store it away from  $\text{CO}_2$ . The *hydrochloride* [2978-59-8] has **m 234°** (from  $\text{EtOH}/\text{Et}_2\text{O}$ ). The *benzoyl* derivative has **m 152-153°** (from  $\text{EtOH}$ ). [Hennion & Teach *J Am Chem Soc* **75** 1653 1953, Hennion & DiGiovanna *J Org Chem* **30** 2645 1965.]

**Methyl n-butyrate** [623-42-7] **M 102.1, b 102.3°/760mm, d<sub>4</sub><sup>20</sup> 0.898, n<sub>D</sub><sup>20</sup> 1.389.** Treat the ester with anhydrous  $\text{CuSO}_4$ , then distil it under dry nitrogen. [*Beilstein* **2** IV 786.]

**S-(+)-2-Methylbutyric acid** [1730-91-2] **M 102.1, b 64°/2mm, 78°/15mm, 90-94°/23mm, 174-175°/atm, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>20</sup> 1.406,  $[\alpha]_{546}^{20} +23^\circ$ ,  $[\alpha]_{\text{D}}^{20} +19.8^\circ$  (neat),  $[\alpha]_{\text{D}}^{13} +18.3^\circ$  (c 6,  $\text{EtOH}$ ), pK<sup>25</sup> 4.76 (for RS).** Purify the acid by distilling it *in vacuo* [Sax & Bergmann *J Am Chem Soc* **77** 1910 1955, Doering & Aschner *J Am Chem Soc* **75** 393 1953]. The *methyl ester* is formed by addition of diazomethane and has **b 112-115°/760mm,  $[\alpha]_{\text{D}}^{27} +21.1^\circ$**  (c 1.7,  $\text{MeOH}$ ). [*Beilstein* **2** IV 888.]

**Methyl carbamate** [598-55-0] **M 75.1, m 54.4-54.8°, 56-58°, b 176-177°/~760mm.** Crystallise the carbamate from \*benzene or distil it. [*Beilstein* **3** H 21.]

**Methyl chloride** [74-87-3] **M 50.5, b -24.1°.** Bubble methyl chloride through a sintered-glass disc dipped into conc  $\text{H}_2\text{SO}_4$ , then wash it with water, condense it at low temperature and fractionally distil it. It has been distilled from  $\text{AlCl}_3$  at  $-80^\circ$ . *Alternatively*, pass it through towers containing  $\text{AlCl}_3$ , soda-lime and  $\text{P}_2\text{O}_5$ , then condense and fractionally distil it. Store it as a gas. [*Beilstein* **1** IV 28.]

**Methyl chloroacetate** [96-34-4] **M 108.5, b 129-130°, d<sub>4</sub><sup>20</sup> 1.230, n<sub>D</sub><sup>20</sup> 1.423.** Shake the ester with saturated aqueous  $\text{Na}_2\text{CO}_3$  (three times), aqueous 50%  $\text{CaCl}_2$  (three times), saturated aqueous  $\text{NaCl}$  (twice), dry ( $\text{Na}_2\text{SO}_4$ ) and fractionally distil it. **Very toxic.** [*Beilstein* **2** IV 480.]

**R-(+)-Methyl 2-chloropropionate** [77287-29-7] **M 122.6, b 49-50°/35mm, 78-80°/120mm, 132-134°/760mm, d<sub>4</sub><sup>20</sup> 1.152, n<sub>D</sub><sup>20</sup> 1.417,  $[\alpha]_{\text{D}}^{20} +26^\circ$  (19.0°) (neat).** Purify the ester by repeated distillation [Walker *J Chem Soc* **67** 916 1895, Walden *Chem Ber* **28** 1293 1985, see also Gless *Synth Commun* **16** 633 1986]. [*Beilstein* **2** H 248.]

**Methyl cyanoacetate** [105-34-0] **M 99.1, f -13°, b 115°/36mm, 200.4-200.9°/761mm, d<sub>4</sub><sup>20</sup> 1.128, n<sub>D</sub><sup>20</sup> 1.420.** Purify the ester by shaking with 10%  $\text{Na}_2\text{CO}_3$  solution, wash well with water, dry with anhydrous  $\text{Na}_2\text{SO}_4$ , and distil it. [*Beilstein* **2** H 584, **2** I 253, **2** II 530, **2** III 1628, **2** IV 1889.]

**Methyl cyanofornate** [17640-15-2] **M 85.1, b 81°/47mm, 97°/751mm, 100-101°/760mm, d<sub>4</sub><sup>20</sup> 1.072, n<sub>D</sub><sup>20</sup> 1.37378.** Purify the ester by fractionation through a 45cm glass helices packed column or a 30cm spinning band column. [Sheppard *J Org Chem* **27** 3756 1962.] It has been distilled through a short Vigreux column, and further purified by recrystallisation from  $\text{Et}_2\text{O}$  at  $-40^\circ$  as white crystals which melt at room temperature.

NMR:  $\delta$  4.0 (CH<sub>3</sub>), and the IR has  $\nu_{\max}$  at 2250 (CN) and 1750 (CO) cm<sup>-1</sup>. [Childes & Weber *J Org Chem* **41** 3486 1976, *Beilstein* **2** III 1587.]

**Methyl decanoate (methyl caprate)** [110-42-9] **M 186.3, b 114°/15mm, 224°/760mm,  $d_4^{20}$  0.874,  $n_D^{20}$  1.426.** Pass the ester through alumina before use and distil in a vacuum. [*Beilstein* **2** IV 1044.]

***N*-Methyldiethanolamine [MDEA, *N,N*-bis(hydroxyethyl)methylamine, 2,2'-methyliminodiethanol]** [105-59-9] **M 119.2, b 75-77°/0.5mm, 115°/5mm, 131°/10mm, 141-142°/18mm, 246-248°/atm,  $d_4^{25}$  1.038,  $n_D^{20}$  1.469,  $pK^{25}$  8.57,  $pK^{35}$  8.31,  $pK^{45}$  8.13,  $pK^{60}$  7.87.** Purify MDEA by fractional distillation preferably under vacuum in a stream of N<sub>2</sub>, and store it under N<sub>2</sub>. The colourless distillate darkens in air and absorbs CO<sub>2</sub> under pressure [Goodridge *Trans Farad Soc* **51** 1703 1955]. It is a tertiary base, hence it does not form an *N*-carbamate salt like primary and secondary amines, but **does** form a *carbonate salt* in the presence of carbonic acid (i.e. CO<sub>2</sub> + H<sub>2</sub>O = H<sub>2</sub>CO<sub>3</sub>). It is soluble in H<sub>2</sub>O and EtOH, but slightly soluble in Et<sub>2</sub>O. The *hydrochloride* is very hygroscopic; the *O,O'*-diacetyl ester has **b 110°/4mm, 133°/15mm, b 251°/atm**, the *O,O'*-bis(4-nitrobenzoyl) ester has **m 112-113°**, the *Reineckate salt* has **m 168°**, the *picrate* crystallises from EtOAc with **m 95-96°**, and with a strong aqueous solution of AuCl<sub>3</sub>/HCl the *chloroaurate salt*, **m 101-102°**, is formed which crystallises from H<sub>2</sub>O. With PtCl<sub>4</sub>/HCl, the *chloroplatinate salt* is obtained, forming orange-yellow rhombic crystals upon recrystallisation from a concentrated aqueous solution which sinter at **145°** and decompose at **148-150°**. The IR spectrum in CCl<sub>4</sub> has C-H str at  $\nu_{\max}$  ( $\epsilon$  mole<sup>-1</sup> L.cm<sup>-1</sup>) 2779 (93), 2844 (81), 2877 (99) and 2948 (107) cm<sup>-1</sup>. [Hanby & Rydon *J Chem Soc* 516 1947, Hill & Meakin *J Chem Soc* 760 1958, Knorr & Matthes *Chem Ber* **31** 1071 1898, *Beilsten* **4** H 284, **4** II 729, **4** III 692, **4** IV 1517.]

**Methyl dodecanoate (methyl laurate)** [111-82-0] **M 214.4, m 5°, b 141°/15mm,  $d_4^{20}$  0.870,  $n_D^{50}$  1.4199.** Pass the ester through alumina before use, and distil it in a vacuum. [*Beilstein* **2** IV 1090.]

***N*-Methyleneaminoacetonitrile (MAAN)** [109-82-0] **M 68.1, m 129°.** Crystallise MAAN from EtOH or acetone. It crystallises nicely from H<sub>2</sub>O but with considerable loss of material. It is an inhibitor of bone growth. [Adams & Langley *Org Synth Coll Vol I* 355 1941.]

**Methyl ether (dimethyl ether)** [115-10-6] **M 46.1, m -141°, b -63.5°/96.5mm, -24°/~760mm,  $d^{25}$  1.918/L (at 1 atmosphere relative to air as 1).** Dry methyl ether by passing over alumina and then BaO, or over CaH<sub>2</sub>, followed by fractional distillation at low temperatures. Its solubility is 37ml per ml of H<sub>2</sub>O at 18°, and it is very soluble in EtOH and Et<sub>2</sub>O. [*Beilstein* **1** IV 1245.]

***N*-Methyl ethylamine hydrochloride** [624-60-2] **M 95.6, m 126-130°,  $pK^{25}$  10.9 (free base).** Crystallise the hydrochloride from absolute EtOH or diethyl ether. Dry it *in vacuo*. [*Beilstein* **4** H 94.]

***N*-Methyl formamide** [123-39-7] **M 59.1, m -3.5°, b 100.5°/25mm,  $d_4^{20}$  1.005.,  $n_D^{52}$  1.4306** Dry it over molecular sieves for 2 days, then distil it under reduced pressure through a column packed with glass helices. Fractionally crystallise it by partial freezing and the solid portion is distilled in a vacuum. [*Beilstein* **4** IV 170.]

**Methyl formate** [107-31-3] **M 60.1, b 31.5°, 34°,  $d_4^{20}$  0.971,  $n_D^{15}$  1.34648,  $n_D^{20}$  1.34332.** Wash the formate with strong aqueous Na<sub>2</sub>CO<sub>3</sub>, dry it with solid Na<sub>2</sub>CO<sub>3</sub> and distil it from P<sub>2</sub>O<sub>5</sub>. (Procedure removes free alcohol or acid.) [*Beilstein* **2** IV 20.]

**2-Methylglutaric acid** [18069-17-5] **M 146.1, m 79°,  $pK_1^{25}$  4.36,  $pK_2^{25}$  5.37.** Crystallise the acid from distilled water, then dry it under vacuum over conc H<sub>2</sub>SO<sub>4</sub>. [*Beilstein* **2** IV 1989.]

**3-Methylglutaric acid** [626-51-7] **M 146.1, m 87°,  $pK_1^{25}$  4.35,  $pK_2^{25}$  5.44.** Crystallise the acid from distilled water, then dry it under vacuum over conc H<sub>2</sub>SO<sub>4</sub>. [*Beilstein* **2** IV 1992.]

**Methylglyoxal** [78-98-8] **M 72.1, b ca 72°/760mm.** Commercial 30% (w/v) aqueous solution is diluted to about 10% and distilled twice, taking the fraction boiling below 50°/20mm Hg. (This treatment does not remove lactic acid.) [*Beilstein* **1** IV 3631.]

**2-Methylhexane** [591-76-4] **M 100.2, b 90.1°**,  $d_4^{20}$  0.678,  $n_D^{20}$  1.38485,  $n_D^{20}$  1.38227. Purify it by azeotropic distillation with MeOH, then wash it with water (to remove the MeOH), dry it over type 4A molecular sieves and distil it. [Beilstein 1 IV 397.]

**3-Methylhexane** [589-34-4] **M 100.2, b 91.9°**,  $d_4^{20}$  0.687,  $n_D^{20}$  1.38864,  $n_D^{20}$  1.38609. Purify it as for 2-methylhexane. [Beilstein 1 IV 399.]

**Methyl hexanoate (methyl caproate)** [106-70-7] **M 130.2, b 52°/15mm, 150°/760mm**,  $d_4^{20}$  0.885,  $n_D^{20}$  1.410. Pass it through alumina and distil it before use. [Beilstein 2 IV 921.]

**Methylhydrazine** [60-34-4] **M 46.1, b 87°/745mm**,  $d_4^{20}$  0.876,  $n_D^{20}$  1.436,  $pK^{30}$  7.87. Dry with BaO, then distil it in a vacuum. Store it under nitrogen. [Beilstein 4 IV 3322.]

**Methyl hydrazinocarboxylate** [6294-89-9] **M 90.1, m 70-73°, b 108°/12mm**. To remove impurities, the material is melted and pumped under vacuum until the vapours are spectroscopically pure [Caminati et al. *J Am Chem Soc* **108** 4364 1986]. Distil it in a vacuum. [Beilstein 3 I 46.]

**Methyl iodide** [74-88-4] **M 141.9, b 42.8°**,  $d_4^{20}$  2.281,  $n_D^{20}$  1.5315. Methyl iodide deteriorates rapidly with liberation of iodine if exposed to light. It is usually purified by shaking it with dilute aqueous  $Na_2S_2O_3$  or  $NaHSO_3$  until colourless, then washing with water, dilute aqueous  $Na_2CO_3$ , and more water, drying with  $CaCl_2$  and distilling. It is stored in a brown bottle away from sunlight in contact with a small amount of mercury, powdered silver or copper. (Prolonged exposure of mercury to methyl iodide forms methylmercuric iodide.) Methyl iodide can be dried further using  $CaSO_4$  or  $P_2O_5$ . An alternative purification is by percolation through a column of silica gel or activated alumina, then distillation. The solution can be degassed by using repeated freeze-pump-thaw cycles. [Beilstein 1 IV 87.]

**O-Methylisourea hydrogen sulfate (2-methylpseudourea sulfate)** [29427-58-5] **M 172.2, m 114-118°, 119°**. Recrystallise the salt from MeOH/Et<sub>2</sub>O (327g of salt dissolved in 1L of MeOH and 2.5L of Et<sub>2</sub>O is added) [Fearing & Fox *J Am Chem Soc* **76** 4382 1954]. The *picrate* has **m 192°** [Odo et al. *J Org Chem* **23** 1319 1958]. [Beilstein 3 IV 143.]

**N-Methyl maleimide** [930-88-1] **M 111.1, m 94-96°**. Crystallise the imide three times from diethyl ether. Dry it *in vacuo*. [Beilstein 21/10 V 5.]

**Methylmalonic acid** [516-05-2] **M 118.1, m 135°(dec)**,  $pK_1^{25}$  3.05,  $pK_2^{25}$  5.76. The acid crystallises as the *hydrate* from water. [Beilstein 2 IV 1932.]

**Methyl methacrylate** [80-62-6] **M 100.1, f -50°, b 46°/100mm, 100°/~760mm**,  $d_4^{20}$  0.937,  $n_D^{20}$  1.4144. Wash the ester twice with aqueous 5% NaOH (to remove inhibitors such as hydroquinone) and twice with water. Dry it with  $CaCl_2$ ,  $Na_2CO_3$ ,  $Na_2SO_4$  or  $MgSO_4$ , then with  $CaH_2$  under nitrogen under reduced pressure. The distillate is stored at low temperatures and redistilled before use. Prior to distilling, inhibitors such as hydroquinone (0.004%),  $\beta$ -naphthylamine (0.2%) or di- $\beta$ -naphthol are sometimes added. Also purify it by boiling with aqueous  $H_3PO_4$  solution and finally with saturated NaCl solution. It is dried for 24 hours over anhydrous  $CaSO_4$ , distilled at 0.1mm Hg at room temperature and stored at -30° [Albeck et al. *J Chem Soc, Faraday Trans 1* **1** 1488 1978]. [Beilstein 2 II 398, 2 III 1279, 2 IV 1519.]

**Methyl methanesulfonate** [66-27-3] **M 110.3, b 59°/0.6mm, 96-98°/19mm**,  $d_4^{20}$  1.300,  $n_D^{20}$  1.4140. Purify the ester by careful fractionation and collecting the middle fraction. Suspected **CARCINOGEN**. Note that  $MeSO_3H$  has **b 167-167.5°/10mm** and methanesulfonic anhydride has **b 138°/10mm** — both are possible impurities. [Beilstein 4 IV 11.]

**Methyl methanethiolsulfonate** [2949-92-0] M 126.2, b 69-71°/0.4mm, 96-97°/4.5mm, 104-105°/10mm, 119°/16mm,  $d_4^{20}$  1.226,  $n_D^{20}$  1.515. Purify it by fractional distillation under reduced pressure, the IR has  $\nu_{\max}$  at 1350, 750  $\text{cm}^{-1}$ . [Applegate et al. *J Org Chem* 38 943 1973, *Beilstein* 4 IV 31.]

**Methyl nitrate** [598-58-3] M 77.0, b 5°/50mm, 65°/760mm,  $d_4^{25}$  1.2322,  $d_4^{15}$  1.2167,  $d_4^{25}$  1.2032. Wash  $\text{MeONO}_2$  once with  $\text{H}_2\text{O}$  then again with  $\text{H}_2\text{O}$  containing a few drops of concentrated NaOH to keep it slightly alkaline (litmus). Dry the ester over anhydrous  $\text{CaCl}_2$ , decant it and use it directly. It is possible to distil it under a vacuum with slow and gentle heating, as a sudden rise in temperature can cause decomposition with copious release of nitrous fumes (use extreme precautions and protection). The middle fraction can then be subjected to several freeze-pump-thaw cycles. [Black & Bakers *Org Synth Col Vol II* 412 1943.] [*Beilstein* 1 H 284, 1 I 141, 1 II 273, 1 III 1201, 1 IV 1254.] The VAPOUR CAN EXPLODE ON HEATING.

**Methyl nitrite** [624-91-9] M 61.0, b  $-18^\circ$ ,  $-17^\circ$ ,  $d_4^{15}$  (liquid) 0.991. Condense  $\text{MeONO}$  in a liquid nitrogen trap. Distil the greenish liquid under vacuum (preferably in a vacuum line), into the first trap containing dry  $\text{Na}_2\text{CO}_3$  to free it from acid impurities then into further  $\text{Na}_2\text{CO}_3$  and fused  $\text{CaCl}_2$  traps before collection at  $-78^\circ$ . It has been distilled through columns that are surrounded by  $\text{Et}_2\text{O}$ /Dri-ice cooled to  $-30^\circ$ . [Leermakers & Ramsperger *J Am Chem Soc* 54 1838 1932, Thompson & Purkis *Trans Farad Soc* 32 675 1936, *Beilstein* 1 H 284, 1 I 141, 1 II 273, 1 III 1201, 1 IV 1253.] **CARCINOGEN.**

**2-Methyl-2-nitro-1,3-propanediol** [77-49-6] M 135.1, m  $145^\circ$ ,  $147-148^\circ$ ,  $149-150^\circ$ . Crystallise it from *n*-butanol or  $\text{Me}_2\text{CO}$  (m  $150.6^\circ$ ). It decomposes on attempted distillation at 10mm. Its solubility in  $\text{H}_2\text{O}$  is 80g/100ml at  $20^\circ$ . [*Beilstein* 1 H 489, 1 II 547, 1 III 2190, 1 IV 2537.]

**2-Methyl-2-nitro-1-propanol** [76-39-1] M 119.1, m  $87-88^\circ$ , b  $94-95^\circ/10\text{mm}$ . Distil it under vacuum and/or crystallise it from petroleum ether or MeOH. [Aistle & Abbott *J Org Chem* 21 1229 1956, Kambe & Yasuda *Bull Soc Chem Jpn* 41 1444 1968, *Beilstein* 1 H 378, 1 III 1546, 1 IV 1604.]

**RS(±)-2-Methyl-2,4-pentanediol (MPD, pinecon, hexylene glycol)** [107-41-5] M 118.2, m  $-40^\circ$ , b  $97^\circ/10\text{mm}$ ,  $107.5-108.5^\circ/25\text{mm}$ ,  $135-136^\circ/40\text{mm}$ ,  $197^\circ/760\text{mm}$ ,  $d_4^{20}$  0.922,  $n_D^{25}$  1.4265. Dry the diol with  $\text{Na}_2\text{SO}_4$ , then  $\text{CaH}_2$ , and fractionally distil it under reduced pressure through a helices packed column, taking precautions to avoid absorption of water. It is soluble in  $\text{H}_2\text{O}$ , EtOH,  $\text{Et}_2\text{O}$ , pentane and hexane. The *diacetate* is a liquid with b  $95^\circ/15\text{mm}$ . It has very harmful vapours; use efficient fume cupboard as it irritates eyes, skin throat and lungs, causing headaches and nausea and is a CNS depressant. [*Beilstein* 1 IV 2565.]

**3-Methyloctane** [2216-33-3] M 128.3, b  $142-144^\circ/760\text{mm}$ ,  $d_4^{20}$  0.719,  $n_D^{20}$  1.407. Take it through a silica gel column and distil it. [Klassen & Ross *J Phys Chem* 91 3668 1987, *Beilstein* 1 IV 455]

**Methyl octanoate (methyl caprylate)** [111-11-5] M 158.2, b  $83^\circ/15\text{mm}$ ,  $193-194^\circ/760\text{mm}$ ,  $d_4^{20}$  0.877,  $n_D^{20}$  1.419. Pass the ester through alumina and distil it before use. [*Beilstein* 2 IV 986.]

**Methyl oleate (methyl *cis*-9-octadecenoate)** [112-62-9]. M 296.5, f  $-19.9^\circ$ , b  $217^\circ/16\text{mm}$ ,  $d_4^{20}$  0.874,  $n_D^{20}$  1.4522. Purify the oleate by fractional distillation under reduced pressure, and by low temperature crystallisation from acetone. Store it in the dark under  $\text{N}_2$ . [*Beilstein* 2 IV 1649.]

**Methylpentane (mixture of isomers).** Pass the mixture through a long column of activated silica gel (or alumina) and collect material that is transparent down to 200nm in the UV.

**2-Methylpentane** [107-83-5] M 86.2, b  $60.3^\circ$ ,  $d_4^{20}$  0.655,  $n_D^{20}$  1.37145,  $n_D^{25}$  1.36873. Purify it by azeotropic distillation with MeOH, followed by washing out the MeOH with water, drying ( $\text{CaCl}_2$ , then sodium), and distilling it. [Forziati et al. *J Res Nat Bur Stand* 36 129 1946.]

**3-Methylpentane** [96-14-0] M 86.2, b  $63.3^\circ$ ,  $d_4^{20}$  0.664,  $n_D^{20}$  1.37652,  $n_D^{25}$  1.37384. Purify it by azeotropic distillation with MeOH, as for 2-methylpentane. Purify it for ultraviolet spectroscopy by passing it through

columns of silica gel or alumina activated by heating for 8 hours at 210° under a stream of nitrogen. Alternatively, treat it with conc (or fuming) H<sub>2</sub>SO<sub>4</sub>, then wash it with water, aqueous 5% NaOH, water again, then dry (CaCl<sub>2</sub>, then sodium), and distil it through a long, glass helices-packed, column. [Beilstein 1 IV 363.]

**2-Methyl-2,4-pentanediol** [107-41-5] M 118.2, b 107.5-108.5°/25mm, d<sub>4</sub><sup>20</sup> 0.922, n<sub>D</sub><sup>25</sup> 1.4265. Dry the diol with Na<sub>2</sub>SO<sub>4</sub>, then CaH<sub>2</sub> and fractionally distil it under reduced pressure through a packed column, taking precautions to avoid absorption of water. [Beilstein 1 IV 2565.]

**2-Methyl-1-pentanol** [105-30-6] M 102.2, b 65-66°/60mm, 146-147°/760mm, d<sub>4</sub><sup>20</sup> 0.827, n<sub>D</sub><sup>20</sup> 1.420. Dry the 1-pentanol with Na<sub>2</sub>SO<sub>4</sub> and distil it. [Beilstein 1 IV 1713.]

**4-Methyl-2-pentanol** [108-11-2] M 102.2, b 131-132°, d<sub>4</sub><sup>20</sup> 0.810, n<sub>D</sub><sup>20</sup> 1.413. Wash the 2-pentanol with aqueous NaHCO<sub>3</sub>, dry and distil it. Purify it further by converting it to the phthalate ester by adding 120ml of dry pyridine and 67g of phthalic anhydride per mole of alcohol, purifying the ester and steam distilling it in the presence of NaOH. The distillate is extracted with ether, and the extract is dried and fractionally distilled. [Levine & Walti *J Biol Chem* 94 367 1931, Beilstein 1 IV 1717.]

**3-Methyl-3-pentanol carbamate (Emylcamate)** [78-28-4] M 145.2, m 56-58.5°. Crystallise the carbamate from 30% aqueous EtOH. [Beilstein 1 IV 1773.]

**4-Methyl-2-pentanone (methyl isobutyl ketone)** [108-10-1] M 100.2, b 115.7°/~760mm d<sub>4</sub><sup>20</sup> 0.801, n<sub>D</sub><sup>20</sup> 1.3958, n<sub>D</sub><sup>25</sup> 1.3938. Reflux the ketone with a little KMnO<sub>4</sub>, wash it with aqueous NaHCO<sub>3</sub>, dry with CaSO<sub>4</sub> and distil it. Acidic impurities are removed by passage through a small column of activated alumina. [Beilstein 1 IV 3305.]

**2-Methyl-1-pentene** [763-29-1] M 84.2, b 61.5-62°, d<sub>4</sub><sup>20</sup> 0.680, n<sub>D</sub><sup>20</sup> 1.395. Water is removed, and formation of peroxides is prevented by several vacuum distillations of 2-methyl-1-pentene from sodium. It is stored with sodium-potassium alloy. [Beilstein 1 IV 841.]

**cis-4-Methyl-2-pentene** [691-38-3] M 84.2, m -134.4°, b 57.7-58.5°, d<sub>4</sub><sup>20</sup> 0.672, n<sub>D</sub><sup>20</sup> 1.388. Dry the *cis*-pentene with CaH<sub>2</sub>, and distil it. [Beilstein 1 IV 841.]

**trans-4-Methyl-2-pentene** [674-76-0] M 84.2, m -140.8°, b 58.5°, d<sub>4</sub><sup>20</sup> 0.669, n<sub>D</sub><sup>20</sup> 1.389. Dry the *trans*-isomer with CaH<sub>2</sub>, and distil it. [Beilstein 1 IV 844.]

**2-Methylpropane-1,2-diamine (1,2-diamino-2-methylpropane)** [811-93-8] M 88.2, b 47-48°/17mm, pK<sub>1</sub><sup>25</sup> 6.25 (6.18), pK<sub>2</sub><sup>25</sup> 9.82 (9.42). Dry the diamine with sodium for 2 days, then distil it from sodium under reduced pressure. [Beilstein 4 IV 1306.]

**2-Methylpropane-1-thiol (isobutylmercaptan)** [513-44-0] M 90.2, b 41.2°/142mm, 88.5°/760mm, n<sub>D</sub><sup>25</sup> 1.43582, pK<sub>Est</sub> ~10.8. Dissolve the thiol in EtOH, and add to 0.25M Pb(OAc)<sub>2</sub> in 50% aqueous EtOH. The precipitated lead mercaptide is filtered off, washed with a little EtOH, and impurities are removed from the molten salt by steam distillation. After cooling, dilute HCl is added dropwise to the residue, and the mercaptan is distilled directly from the flask. Water is separated from the distillate, and the mercaptan is dried (Na<sub>2</sub>CO<sub>3</sub>) and distilled under nitrogen. [Mathias *J Am Chem Soc* 72 1897 1950, Beilstein 1 H 378, 1 I 191, 1 II 412, 1 III 1565, 1 IV 1605.]

**2-Methylpropane-2-thiol (tert-butylmercaptan)** [75-66-1] M 90.2, b 61.6°/701mm, 66°/760mm, d<sub>4</sub><sup>25</sup> 0.79426, n<sub>D</sub><sup>25</sup> 1.41984, pK<sup>25</sup> 11.22. Dry the thiol for several days over CaO, then distil it from CaO. Purify it as for 2-methylpropane-1-thiol above. [Beilstein 1 H 383, 1 II 416, 1 III 1589, 1 IV 1634.]

**2-Methyl-1-propanol (isobutanol)** [78-83-1] M 74.1, b 107.9°/760mm, d<sub>4</sub><sup>20</sup> 0.804, n<sub>D</sub><sup>15</sup> 1.39768, n<sub>D</sub><sup>25</sup> 1.3939. Isobutanol is dried by refluxing with CaO and BaO for several hours, followed by treatment with calcium or aluminium amalgam, then fractional distilling it from sulfanilic or tartaric acids. More exhaustive purifications

involve formation of phthalate or borate esters. Heating it with phthalic anhydride gives the *acid phthalate* which, after crystallisation to constant melting point (*m* 65°) from petroleum ether, is hydrolysed with aqueous 15% KOH. The alcohol is distilled off as the water azeotrope and dried (K<sub>2</sub>CO<sub>3</sub>, then anhydrous CuSO<sub>4</sub>), and finally magnesium turnings, followed by fractional distillation. [Hüchel & Ackermann *J Prakt Chem* **136** 15 1933.] The borate ester is formed by heating the dried alcohol for 6 hours in an autoclave at 160-175° with a quarter of its weight of boric acid. After fractional distillation under vacuum, the ester is hydrolysed by heating for a short time with aqueous alkali and the alcohol is dried with CaO and distilled. [Michael et al. *J Am Chem Soc* **38** 653 1916.] Alternatively, dry the alcohol with K<sub>2</sub>CO<sub>3</sub>, CaSO<sub>4</sub> or CaCl<sub>2</sub>, filter and fractionally distil it. For further drying, the redistilled alcohol can be refluxed with the appropriate alkyl phthalate or succinate as described under *ethanol*. [Beilstein **1** IV 1588.]

**Methyl propiolate** [922-67-8] **M 84.1, b 100°/atm, 102°/atm, 103-105°/atm, d<sub>4</sub><sup>20</sup> 0.945, n<sub>D</sub><sup>20</sup> 1.4080.** Purify the propiolate by fractional distillation and collecting the middle fraction, note that propiolic acid has a higher boiling point [144°(dec)/760mm]. [Beilstein **2** IV 1688.] **LACHRYMATORY.**

**N-Methylpropionamide** [1187-58-2] **M 87.1, f -30.9°, b 103°/12-13mm, d<sub>4</sub><sup>20</sup> 0.934, n<sub>D</sub><sup>25</sup> 1.4356.** The amide is a colourless, odourless, neutral liquid at room temperature with a high dielectric constant. The amount of water present can be determined directly by Karl Fischer titration, GLC and NMR have been used to detect unreacted propionic acid. Commercial material of high quality is available, probably from the condensation of anhydrous methylamine with 50% excess of propionic acid. Rapid heating to 120-140° with stirring favours the reaction by removing water either directly or as the ternary xylene azeotrope. The quality of the distillate improves during the distillation.

*N*-Methylpropionamide can be dried over CaO. Water and unreacted propionic acid are removed as their xylene azeotropes. It is then distilled in a vacuum. Material used as an electrolyte solvent (specific conductance less than 10<sup>-6</sup> ohm<sup>-1</sup> cm<sup>-1</sup>) is obtained by fractional distillation under reduced pressure, and storage over BaO or molecular sieves because it readily absorbs moisture from the atmosphere on prolonged storage. [Hoover *Pure Appl Chem* **37** 581 1974, *Recommended Methods for Purification of Solvents and Tests for Impurities*, Coetzee Ed., Pergamon Press, 1982, Beilstein **4** IV 183.]

**Methyl propionate** [554-12-1] **M 88.1, b 79.7°.** Wash the ester with saturated aqueous NaCl, then dry it with Na<sub>2</sub>CO<sub>3</sub> and distil it from P<sub>2</sub>O<sub>5</sub>. (This removes any free acid and alcohol.) It has also been dried with anhydrous CuSO<sub>4</sub>. [Beilstein **2** IV 104.]

**Methyl *n*-propyl ether** [557-17-5] **M 74.1, b 39°/743mm d<sub>4</sub><sup>20</sup> 0.736, n<sub>D</sub><sup>14</sup> 1.3602, pK<sup>25</sup> -3.79 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Dry it with CaSO<sub>4</sub>, then pass the ether through a column of alumina (to remove peroxides) and fractionally distil it. [Beilstein **1** H 354, **1** I 178, **1** II 367, **1** III 1413, **1** IV 1421.]

**Methyl *n*-propyl ketone (pentan-2-one)** [107-87-9] **M 86.1, b 102.4°, d<sub>4</sub><sup>20</sup> 0.807, n<sub>D</sub><sup>20</sup> 1.3903.** Purify the ketone by refluxing it with a little KMnO<sub>4</sub>, dry it with CaSO<sub>4</sub> and distil it. It can be converted to its bisulfite addition compound by shaking with excess saturated aqueous NaHSO<sub>3</sub> at room temperature, cooling to 0°, filtering, washing with diethyl ether and drying. Steam distillation of the adduct gives a distillate from which the ketone is recovered, washed with aqueous NaHCO<sub>3</sub> and distilled water, dried (K<sub>2</sub>CO<sub>3</sub>) and fractionally distilled. [Waring & Garik *J Am Chem Soc* **78** 5198 1956, Beilstein **1** IV 3271.]

**(±)-3-Methyl-1-propyn-3-ol carbamate (Meparfynol carbamate)** [302-66-9] **M 141.2, m 55.8-57°, 56-58°, 120-121°/16mm.** Crystallise it from \*C<sub>6</sub>H<sub>6</sub>, hexane, ether/petroleum ether or cyclohexane. [Beilstein **1** IV 65.] It is a sedative.

**Methyl stearate** [122-61-8] **M 298.5, m 41-43°, b 181-182°/4mm.** Crystallise the ester from petroleum ether or distil it in a vacuum. [Beilstein **2** IV 1216.]

**Methylsuccinic acid** [498-21-5] **M 132.1, m 115.0°, pK<sub>1</sub><sup>25</sup> 3.88, pK<sub>2</sub><sup>25</sup> 5.35.** Crystallise the acid from water. [Beilstein **2** IV 1948.]

**N-Methylthioacetamide** [5310-10-1] **M 89.1, m 59°**. Recrystallise the amide from \*benzene or EtOH. [Todd et al. *Chem Ber* **69** 220 1936, *Beilstein* **4** I 329, **4** III 124.]

**Methyl trifluoromethanesulfonate (methyl triflate)** [333-27-7] **M 164.1, b 97-97.5°/736mm, 99°/~760mm, 100-102°/~760mm,  $d_4^{20}$  1.496,  $n_D^{25}$  1.3238**. It is a strong methylating agent but is corrosive and **POISONOUS**. Fractionate it carefully and collecting the middle fraction (use an efficient fume cupboard) and keep away from moisture. It is a **POWERFUL ALKYLATING AGENT** and a strong **IRRITANT**. [IR: Gramstad & Haszeldine *J Chem Soc* 173 1956, *J Chem Soc* 4069 1957.] *Trifluoromethanesulfonic acid* (triflic acid) [1493-13-6] **M 151.1, boils higher (b 162°/atm), has a pKa of 3.10, and is TOXIC and hygroscopic**. [Hansen *J Org Chem* **30** 4322 1965, Kurz & El-Nasr *J Am Chem Soc* **104** 5823 1982, *Beilstein* **3** IV 34.]

**Methyl vinyl ketone (3-buten-2-one)** [78-94-4] **M 70.1, b 62-68°/400mm, 79-80°/760mm,  $d_4^{20}$  0.845,  $n_D^{20}$  1.413**. It forms an 85% azeotrope with water. After drying with  $K_2CO_3$  and  $CaCl_2$  (with cooling), the ketone is distilled at low pressures. [*Beilstein* **1** IV 3444.]

**Methyl vinyl sulfone** [3680-02-2] **M 106.1, b 116-118°/20mm,  $d_4^{20}$  1.215,  $n_D^{20}$  1.461**. Pass the sulfone through a column of alumina, then de-gas, distil it in a vacuum line and store it at  $-190^\circ$  until required. [*Beilstein* **1** III 1866.]

**N-Monobutyl urea** [592-31-4] **M 116.2, m 96-98°,  $pK_{Est}$  ~0.2**. Crystallise the urea from EtOH/water, then dry it under vacuum at room temperature. [*Beilstein* **4** I 371, **4** IV 578.]

**N-Monoethyl urea** [625-52-5] **M 88.1, m 92-95°,  $pK_{Est}$  ~0.2**. Crystallise the urea from EtOH/water, then dry it under vacuum at room temperature. [*Beilstein* **4** IV 369.]

**N-Monomethyl urea** [598-50-5] **M 74.1, m 93-95°,  $pK_{Est}$  ~0.2**. Crystallise the urea from EtOH/water, then dry it under vacuum at room temperature. [*Beilstein* **4** IV 205.]

**N-Monopropyl urea** [627-06-5] **M 102.1, m 107°, 110°,  $pK_{Est}$  ~0.2**. Crystallise the urea from EtOH or EtOH/Et<sub>2</sub>O. [Biovin & Biovin *Can J Chem* **29** 479 1951, IR: Biovin & Biovin *Can J Chem* **32** 563 1954, *Beilstein* **4** H 142, **4** III 261, **4** IV 482.]

**Mucochloric acid (2,3-dichloro-4-oxo-2-butenic acid)** [87-56-9] **M 169.0, m 124-126°,  $pK^{25}$  4.20**. Crystallise the acid twice from water (charcoal). [*Beilstein* **3** IV 1720.]

**trans, trans-Muconic acid (hexa-2,4-dienedioic acid)** [3588-17-8] **M 142.1, m 300°,  $pK^{25}$  4.51, for cis, cis  $pK^{25}$  4.49**. Crystallise the diacid from H<sub>2</sub>O. [*Beilstein* **2** IV 2298.]

**Myristic acid (tetradecanoic acid)** [544-63-8] **M 228.4, m 58°,  $pK^{20}$  6.3 (50% aqueous EtOH),  $pK_{Est}$  ~4.9 (H<sub>2</sub>O)**. Purify the acid via the *methyl ester* (b 153-154°/10mm,  $n^{25}$  1.4350), as for capric acid. [Trachtman & Miller *J Am Chem Soc* **84** 4828 1962.] Also purify it by zone melting. It crystallises from petroleum ether, and is dried in a vacuum desiccator containing shredded wax. [*Beilstein* **2** IV 1126.]

**Neopentane (2,2-dimethylpropane)** [463-82-1] **M 72.2, flash point 79.3°, m -19.8°, b 9.5°/760mm,  $d_4^{20}$  0.6737,  $n_D^{20}$  1.38273**. It is freed from isobutene by passage over conc H<sub>2</sub>SO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub>, and through silica gel. [*Beilstein* **1** H 141, **1** I 50, **1** II 104, **1** 369, **1** IV 333.]

**Nerolidol (3,7,11-trimethyl-1,6,10-dodecatrien-3-ol)** **M 222.4 [cis/trans 7212-44-4] b 122°/3mm,  $d_4^{20}$  0.73,  $n_D^{20}$  1.477, [cis 3790-78-1] b 70°/0.1mm, [trans 40716-66-3] b 78°/0.2mm, 145-146°/2mm**. Purify it by TLC on plates of Kieselguhr G [McSweeney *J Chromatogr* **17** 183 1965] or silica gel impregnated with AgNO<sub>3</sub>, using 1,2-CH<sub>2</sub>Cl<sub>2</sub>/CHCl<sub>3</sub>/EtOAc/PrOH (10:10:1:1) as solvent system. Also by GLC on butanediol succinate (20%) on Chromosorb W. Stored it under N<sub>2</sub> at ~5° in the dark. [*Beilstein* **1** IV 2336.]

**Nitrilotriacetic acid** [tris(carboxymethyl)amine, NTA, Complexone 1] [139-13-9] **M 191.1, m 247°(dec), pK<sub>1</sub> 0.8, pK<sub>2</sub> 1.71, pK<sub>3</sub> 2.47, pK<sub>4</sub> 9.71.** Crystallise it from water and dry it at 110°. [Beilstein 4 IV 2441.]

**Nitroethane** [79-24-3] **M 75.1, b 115°, d<sub>4</sub><sup>20</sup> 1.049, n<sub>D</sub><sup>20</sup> 1.3920, n<sub>D</sub><sup>25</sup> 1.39015, pK<sup>25</sup> 8.60 (8.46, pH equilibrium requires ca 5 minutes).** Purify it as described for *nitromethane* below. A spectroscopic impurity can be removed by shaking it with activated alumina, decanting and distilling it rapidly. [Beilstein 1 IV 170.]

**Nitroguanidine** [556-88-7] **M 104.1, m 246-246.5°(dec), 257°, pK<sub>1</sub><sup>25</sup> -0.55, pK<sub>2</sub><sup>25</sup> 12.20.** Crystallise it from water (20ml/g). The *nitrate* has **m 147°(dec)**(prisms, H<sub>2</sub>O). [Beilstein 3 H 126, 3 III 236.]

**Nitromethane** [75-52-5] **M 61.0, f -28.5°, b 101.3°, d<sub>4</sub><sup>20</sup> 1.13749, d<sub>30</sub><sup>20</sup> 1.12398, n<sub>D</sub><sup>20</sup> 1.3819, n<sub>D</sub><sup>30</sup> 1.37730, pK<sup>25</sup> 10.21.** Nitromethane is generally manufactured by gas-phase nitration of methane. The usual impurities include aldehydes, nitroethane, water and small amounts of alcohols. Most of these can be removed by drying with CaCl<sub>2</sub> or by distillation to remove the water/nitromethane azeotrope, followed by drying with CaSO<sub>4</sub>. Phosphorus pentoxide is not suitable as a drying agent. [Wright et al. *J Chem Soc* 199 1936.] The purified material should be stored by dark bottles, away from strong light, in a cool place. Purifications using extraction are commonly used. For example, Van Looy and Hammett [*J Am Chem Soc* 81 3872 1959] mixed about 150ml of conc H<sub>2</sub>SO<sub>4</sub> with 1L of nitromethane and allowed it to stand for 1 or 2 days. The solvent was washed with water, aqueous Na<sub>2</sub>CO<sub>3</sub>, and again with water, then dried for several days with MgSO<sub>4</sub>, filtered again with CaSO<sub>4</sub>. It was fractionally distilled before use. Smith, Fainberg and Winstein [*J Am Chem Soc* 83 618 1961] washed it successively with aqueous NaHCO<sub>3</sub>, aqueous NaHSO<sub>3</sub>, water, 5% H<sub>2</sub>SO<sub>4</sub>, water and dilute NaHCO<sub>3</sub>. The solvent was dried with CaSO<sub>4</sub>, then percolated through a column of Linde type 4A molecular sieves, followed by distillation from some of this material (in powdered form). Buffagni and Dunn [*J Chem Soc* 5105 1961] refluxed it for 24 hours with activated charcoal while bubbling a stream of nitrogen through the liquid. The suspension was filtered, dried (Na<sub>2</sub>SO<sub>4</sub>) and distilled, then passed through an alumina column and re-distilled. It has also been refluxed over CaH<sub>2</sub>, distilled and kept under argon over 4A molecular sieves.

It has been purified by zone melting at low temperature, or by distillation under vacuum at 0°, subjecting the middle fraction to several freeze-pump-thaw cycles. An impure sample containing higher nitroalkanes and traces of cyanoalkanes was purified (on the basis of its NMR spectrum) by crystallisation from diethyl ether at -60° (cooling in Dry-ice) [Parrett & Sun *J Chem Educ* 54 448 1977].

Fractional crystallisation is more effective than fractional distillation from Drierite in purifying nitromethane for conductivity measurements. [Coetzee & Cunningham *J Am Chem Soc* 87 2529 1965.] Specific conductivities around 5 x 10<sup>-9</sup> ohm<sup>-1</sup>cm<sup>-1</sup> were obtained. [Beilstein 1 IV 100.]

**1-Nitropropane** [108-03-2] **M 89.1, b 131.4°, d<sub>4</sub><sup>20</sup> 1.004, n<sub>D</sub><sup>20</sup> 1.40161, n<sub>D</sub><sup>25</sup> 1.39936, pK<sup>25</sup> 8.98.** Purify it as for *nitromethane*. [Beilstein 1 IV 229.]

**2-Nitropropane** [79-46-9] **M 89.1, b 120.3°, d<sub>4</sub><sup>20</sup> 0.989, n<sub>D</sub><sup>20</sup> 1.3949, n<sub>D</sub><sup>25</sup> 1.39206, pK<sup>25</sup> 7.68.** Purify it as for *nitromethane*. [Beilstein 1 IV 230.]

**N-Nitrosodiethanolamine (NDELA)** [1116-54-7] **M 134.4, b 100°/2.6x10<sup>-5</sup>mm, 125°/0.01mm, n<sub>D</sub><sup>20</sup> 1.4849.** Purify NDELA by dissolving the amine (0.5g) in 1-propanol (10ml) and 5g of anhydrous Na<sub>2</sub>SO<sub>4</sub> added with stirring. After standing for 1-2 hours, it is filtered and passed through a chromatographic column packed with 10ml of AG 50W x 8 (H<sup>+</sup> form 50-100mesh, a strongly acidic cation exchanger). The eluent and washings (50 ml EtOH) are combined and evaporated to dryness at 35°. It has also been extracted with EtOH from the nitrosation mixture of ethanolamine, filtered and distilled under high vacuum. [Fukuda et al. *Anal Chem* 53 2000 1981, Jones & Wilson *J Chem Soc* 550, 1949, *Beilstein* 1 III 721, see Spiegelhalder et al. *N-Nitroso Compounds: Occurrence Biological Effects and Relevance in Human Cancer* (eds. O'Neill et al. IARC Scientific Publications No 57; IARC Lyon p943 1984.) **Possible CARCINOGEN.**

**Nitrourea** [556-89-8] **M 105.1, m 158.4-158.8°(dec).** Crystallise it from EtOH/petroleum ether. Dry it *in vacuo* ~50°. [Ingersoll & Arenendt *Org Synth Coll Vol I* 417 1941.]

**n-Nonane** [111-84-2] **M 126.3, b 150.8°, d<sub>4</sub><sup>20</sup> 0.719, n<sub>D</sub><sup>20</sup> 1.40542, n<sub>D</sub><sup>25</sup> 1.40311.** Fractionally distil *n*-nonane,

then stir it with successive volumes of conc  $\text{H}_2\text{SO}_4$  for 12 hours each until no further coloration is observed in the acid layer. Then wash it with water, dry with  $\text{MgSO}_4$  and fractionally distil it. *Alternatively*, it is purified by azeotropic distillation with 2-ethoxyethanol, followed by washing out the alcohol with water, drying and distilling it. [Forziati et al. *J Res Nat Bur Stand* **36** 129 1946, *Beilstein* **1** IV 447.]

**Nylon powder.** Pellets are purified by dissolving them in ethylene glycol under reflux, then precipitating nylon as a white powder by adding EtOH at 25°. This is washed with EtOH and dried at 100° under vacuum.

***n*-Octacosane** [630-02-4] **M 394.8, m 62.5°.** Purify it by forming its adduct with urea, washing it and crystallising it from acetone/water. [McCubbin *Trans Faraday Soc* **58** 2307 1962.] Crystallise it then from hot filtered isopropyl ether solution (10ml/g). [*Beilstein* **1** IV 588.]

***n*-Octacosanol (octacosyl alcohol)** [557-61-9] **M 410.8, m 83.4°, 84°.** Recrystallise it from large volumes of  $\text{Me}_2\text{CO}$ . Sublime it at 200-250°/1mm instead of distilling it. [*Beilstein* **2** IV 1318.]

***n*-Octadecane** [593-45-3] **M 254.5, m 28.1°, b 173.5°/10mm, 316.1°/760mm,  $d_4^{20}$  0.7768,  $n_D^{20}$  1.4390.** Crystallise it from acetone and distil it from sodium in a vacuum. [*Beilstein* **1** IV 553.]

**Octadecyl acetate** [822-23-1] **M 312.5, m 32.6°, b 166-168°/1mm, 172-174°/1.5mm.** Distil the ester under high vacuum, then crystallise it from  $\text{Et}_2\text{O}/\text{MeOH}$ , EtOH (**m 32.8°**) or  $\text{Me}_2\text{CO}$  (**m 32.4°**). Also recorded are **m 30.2°**; and **35.3°** for an  $\alpha$  form as well as **38-39°** for a  $\beta$  form. [Phillips & Mumford *J Chem Soc* 1663 1934, *Beilstein* **2** H 136, **2** II 147, **2** III 266, **2** IV 171.]

***n*-Octadecyl alcohol (stearyl alcohol, octadecanol)** [112-92-5] **M 270.5, m 61°, b 153-154°/0.3mm.** Crystallise octadecanol from MeOH, or dry  $\text{Et}_2\text{O}$  and  $^*\text{C}_6\text{H}_6$ , then fractionally distil it *in vacuo*. Also purify it by column chromatography. Free it from cetyl alcohol by zone refining. [*Beilstein* **1** IV 1888.]

***n*-Octadecylamine (1-aminooctadecane, stearylamine, ODA)** [124-30-1] **M 269.5, m 49-52°, 50-52°, 52-54°, b 183.0-183.1°/5mm, 232°/32mm,  $\text{pK}^{25}$  10.60.** *n*-Octadecylamine can be prepared from nonadecanoic acid (*n*-nonadecylic acid, [646-30-0]) by Schmidt's hydrazoic acid method whereby the carboxylic acid (15.8g, 53mmol) in  $^*\text{C}_6\text{H}_6$  (500ml) is treated carefully with concentrated  $\text{H}_2\text{SO}_4$  (30ml), stirred vigorously at 40° and hydrazoic acid (52ml of a 5.3% solution in  $^*\text{C}_6\text{H}_6$ , 1.2 equivalents, i.e. 63.3mmol; see [7782-79-8], **POISONOUS** use efficient fumehood) is added slowly. After evolution of  $\text{N}_2$  and  $\text{CO}_2$  have ceased (~2 hours) the acid layer is poured into cold  $\text{H}_2\text{O}$  whereby octadecylamine sulfate precipitates (~96% yield decomposing at ~200°). [see Briggs et al *J Chem Soc* 61 1942.] The free base is obtained by shaking with 10% KOH solution, extracting into  $\text{Et}_2\text{O}$ , drying ( $\text{K}_2\text{CO}_3$ ), evaporating and distilling the residue in a vacuum. Its FT-IR (film) has  $\nu_{\text{max}}$  at 2923.9, 1466.3, 792.4, 720.9  $\text{cm}^{-1}$ ; its  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 0.89 (t,  $J = 7$  Hz, 3H,  $\text{CH}_3$ ), 1.01 (s), 1.25 (s, 10H), 1.43 (m, 2H), 2.67 (t,  $J = \sim 7$  Hz, 2H,  $\text{CH}_2\text{-N}$ ); and the  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ ,  $\text{CHCl}_3$  standard) has  $\delta$  at 42.35, 33.99, 31.95, 29.72, 29.56, 29.39, 26.94, 22.71, 14.11. It has also been prepared by the Gabriel Synthesis whereby octadecanol is converted to octadecyl iodide, m 29-32° (by reaction with  $\text{P}_2\text{O}_5/\text{KI}$ ), then with potassium phthalimide in DMF, to the *N-n*-octadecylphthalimide, m 80-81°, followed by hydrazinolysis with 85%  $\text{NH}_2\text{NH}_2 \cdot \text{H}_2\text{O}$ , and isolation as the *hydrochloride* [1838-08-0] **m 162-163°** after crystallisation from EtOH/ $\text{Et}_2\text{O}$  [Wood *J Chem Soc* 3327 1953]. Another method of preparation include reduction of stearic amide with LAH in ~60% yield using a Soxhlet extractor with the amide in the thimble and LAH/ $\text{Et}_2\text{O}$  in the boiling flask [Murr & Lester *J Am Chem Soc* 77 1684 1955]. The *N-acetyl derivative* has **m 84-85°** (EtOH/ $\text{Et}_2\text{O}$ ), and the *N-methyl derivative* [2439-55-6] has **m 42-46°, b 155°/0.5mm**. [*Beilstein* **4** II 661, **4** III 431, **4** IV 825.] The base and its derivatives are **skin and eye irritants**, and release histamine.

**Octadecyl ether (dioctadecyl ether)** [6297-03-6] **M 523.0, m 59.4°,  $n_D^{60}$  1.440.** Distil the ether in a vacuum, then crystallise it from MeOH/ $^*\text{C}_6\text{H}_6$ , MeOH (**m 58.5-59.5°**) or  $\text{Me}_2\text{CO}$  (**m 59.5°**). It has an  $\alpha$  form **m 57.8°** and a  $\beta$  form **m 40°**. [*Beilstein* **1** IV 1891.]

**Octafluoropropane (profluorane)** [76-19-7] **M 188.0, m -183°, b -38°**. Purify it for pyrolysis studies by passage through a copper vessel containing  $\text{CoF}_3$  at about 270°, then fractionally distil it. [Steunenberg & Cady *J Am Chem Soc* **74** 4165 1952.] Also purify it by several trap-to-trap distillations at low temperatures [Simons & Block *J Am Chem Soc* **59** 1407 1937].

**n-Octane** [111-65-9] **M 114.2, m -56.8°, b 19.2°/10mm, 125.6°/760mm,  $d_4^{20}$  0.704,  $n_D^{20}$  1.39743,  $n_D^{25}$  1.39505**. Extract the octane repeatedly with conc  $\text{H}_2\text{SO}_4$  or chlorosulfonic acid, then wash it with water, dry and distil it. Alternatively, purify it by azeotropic distillation with EtOH, followed by washing with water to remove the EtOH, drying and distilling it. For further details, see *n-heptane*. It is also purified by zone melting at low temperature. [Beilstein **1 H 159, 1 I 60, 1 II 122, 1 III 457, 1 IV 412.**]

**RS-(±)-Octane-1,2-diol** [1117-86-8] **M 146.2, m 30-30.5°, 36°, b 103-105°/0.5mm, 131-132°/10mm**. Distil the diol *in vacuo* and/or recrystallise it from petroleum ether. The  $\alpha$ -*naphthylurethane* has **m 112-114°**. [Beilstein **1 III 2217, 1 IV 2590.**] *S*-(-)-*Octane-1,2-diol* [87720-91-0] also crystallises from petroleum ether with **m 35-37°** and  $[\alpha]_D^{17}$  -4.7° (c 35, EtOH) [Späth et al. *Chem Ber* **66** 598 1933]; *R*-(+)-*octane-1,2-diol* [87720-90-9] has similar properties but with a positive optical rotation.

**Octane-1,8-diol (octamethylene glycol)** [629-41-4] **M 146.2, m 59-61°, b 172°/20mm**. Recrystallise the diol from EtOH and distil it in a vacuum. [Beilstein **1 IV 2592.**]

**1-Octanethiol** [111-88-6] **M 146.3, b 86°/15mm, 197-200°/760mm,  $d_4^{20}$  0.8433,  $n_D^{20}$  1.4540,  $pK^{25}$  10.72**(dilute *t*-BuOH). Pass the thiol through a column of alumina and work under  $\text{N}_2$ , or Ar. Distil it under  $\text{N}_2$  and a vacuum. Store it under  $\text{N}_2$ , or Ar in the dark. [Battacharyya et al. *J Chem Soc, Faraday Trans 1* **82** 135 1986, Fletcher *J Am Chem Soc* **68** 2727 1946]. [Beilstein **1 III 1710, 1 IV 1767.**]

**1-Octene** [111-66-0] **M 112.2, b 121°/742mm,  $d_4^{20}$  0.716,  $n_D^{20}$  1.4087**. Distil 1-octene under nitrogen from sodium which removes water and peroxides. Peroxides can also be removed by percolation through dried, acid washed, alumina. Store it under  $\text{N}_2$ , or Ar in the dark. [Strukul & Michelin *J Am Chem Soc* **107** 7563 1985, Beilstein **1 H 221, 1 II 199, 1 IV 874.**]

**trans-2-Octene** [13389-42-9] **M 112.2, b 124-124.5°/760mm,  $d_4^{20}$  0.722,  $n_D^{20}$  1.4132**. Purify it as for 1-octene above. [Beilstein **1 IV 879.**]

**n-Octyl alcohol** [111-87-5] **M 130.2, b 98°/19mm, 195.3°/760mm,  $d_4^{20}$  0.828,  $n_D^{20}$  1.43018**. Fractionally distil it under reduced pressure. Dry it with sodium and again fractionally distil or reflux with boric anhydride and re-distilled (**b 195-205°/5mm**), the distillate being neutralised with NaOH and again fractionally distilled. Also purify it by distillation from Raney nickel and by preparative GLC. [Beilstein **1 IV 1756.**]

**n-Octylamine (1-aminooctane, caprylamine)** [111-86-4] **M 129.2, m -5-1°, b 62-64°/11mm 175-177°/745mm, 185-187°/atm,  $d_4^{20}$  0.7819,  $d_4^{25}$  0.728,  $n_D^{20}$  1.4292,  $pK^{25}$  10.57**. *n*-Octylamine has been prepared from nonanoic acid (pelargonic acid, see [112-05-0]) by Schmidt's hydrazoic acid method as described for octadecylamine above. [Adamson & Kenner *J Chem Soc* 383 1943.] The syntheses described below for octadecylamine can be used for preparing octylamine. A more recent preparation involves formation of tri-*n*-octylborane from 1-octene and  $\text{BH}_3$ .THF *in situ*, then reaction with  $\text{NaN}_3/\text{HCl}$  to give a 79% yield of octylamine [Kabalka et al. *Organometallics* **6** 1369 1987]. The free base can be isolated from its salts or the picrate by treatment with strong aqueous alkaline solution, extraction into  $\text{Et}_2\text{O}$  or  $\text{CH}_2\text{Cl}_2$ , drying ( $\text{K}_2\text{CO}_3$ ), evaporating and distilling the residue preferably in a vacuum. Solubility in  $\text{H}_2\text{O}$  is ~2%. It is a strong base which readily absorbs  $\text{CO}_2$  in moist air and should be stored under  $\text{N}_2$ . Its FT-IR (film) has  $\nu_{\text{max}}$  at 3371.9, 2924.6, 1617.3, 1467.0, 1378.3, 1072.7, 822.1, 722.9  $\text{cm}^{-1}$ ; its  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 0.89 (t,  $J = \sim 7$  Hz, 3H,  $\text{CH}_3$ ), 1.15 (s), 1.29 (s, 6 $\text{CH}_2$ ), 1.42 (t,  $J = \sim 7$  Hz,  $\text{NH}_2$ ), 2.69 (t,  $J = \sim 7$  Hz, 2H,  $\text{CH}_2\text{-N}$ ) ppm; and the  $^{13}\text{C}$  NMR (300MHz,  $\text{CDCl}_3$ ,  $\text{CHCl}_3$  standard) has  $\delta$  at 42.32, 33.96, 31.88, 29.51, 29.34, 26.94, 22.69, 14.10 ppm. The picrate [78498-55-2] crystallises as plates with **m 111.5-112.5°**, the *N*-acetyl derivative [7462-62-6] has **b 148-149°/3mm**, and the *N*-methyl derivative [2439-54-5] has **b 60-65°/3mm, 75-78°/13mm**. [Beilstein **4 H 196, 4 I 386, 4 II 655, 4 III 379, 4 IV 751.**] *n*-Octylamine is quite a strong primary base capable of forming a carbamate

salt of the type  $\text{RNHCO}_2^- \text{RNH}_3^+$  in the presence of  $\text{CO}_2$  and  $\text{H}_2\text{O}$  (see storage above). The free base and its derivatives are **skin and eye irritants**.

***n*-Octylammonium hexadecanoate** [88020-97-7] **M 385.7, m 52-53°**. Purify it by several recrystallisations from *n*-hexane or ethyl acetate. The solid is then washed with cold anhydrous diethyl ether, and dried *in vacuo* over  $\text{P}_2\text{O}_5$ . [Beilstein 4 IV 751 for octylamine.]

***n*-Octylammonium octadecanoate** [32580-92-0] **M 413.7, m 56-57°**. Purify it as for the *hexadecanoate* above.

***n*-Octylammonium tetradecanoate** [17463-35-3] **M 358.6, m 46-48°**. Purify it as the *hexadecanoate* above.

***n*-Octyl bromide** [111-83-1] **M 193.1, b 201.5°,  $d_4^{20}$  1.118,  $n_D^{25}$  1.4503**. Shake the bromide with  $\text{H}_2\text{SO}_4$ , wash it with water, dry with  $\text{K}_2\text{CO}_3$  and fractionally distil it. [Beilstein 1 IV 422.]

**1-Octyne** [629-05-0] **M 110.2, b 76-77°/150mm, 126.2°/760mm,  $d_4^{20}$  0.717,  $n_D^{25}$  1.4159**. Distil 1-octyne from  $\text{NaBH}_4$  to remove peroxides. Fractionate it through a 10inch Widmer column at 125-126°/759mm [Slettinger & Dawson *J Org Chem* 14 853 1949.] [Beilstein 1 III 1005, 1 IV 1034.]

**Oleic acid (cis-9-octadecenoic acid, olainic acid)** [112-80-1] **M 282.5, m 16°, b 145°/0.1mm, 194-195°/1.2mm, 228-229°/15mm, 360°(dec),  $d_4^{20}$  0.891,  $n_D^{30}$  1.4571,  $pK_{25}^{25}$  6.42 (50% aqueous EtOH),  $pK_{\text{Est}} \sim 4.8$  ( $\text{H}_2\text{O}$ )**. Purify the acid by fractional crystallisation from its melt, followed by molecular distillation at  $10^{-3}$ mm, or by conversion to its methyl ester, the free acid can be crystallised from acetone at  $-40^\circ$  to  $-45^\circ$  (12ml/g). For purification by the use of lead and lithium salts see Keffler and McLean [*J Soc Chem Ind (London)* 54 176T 1935]. Purification based on direct crystallisation from acetone is described by Brown and Shinowara [*J Am Chem Soc* 59 6 1937, *pK White J Am Chem Soc* 72 1857 1950]. [Beilstein 2 H 463, 2 I 198, 2 II 429, 2 III 1387, 2 IV 1641.]

**Oleyl alcohol** [143-28-2] **M 268.5, b 182-184°/1.5mm,  $d_4^{20}$  0.847,  $n_D^{27.5}$  1.4582**. Purify it by fractional crystallisation at  $-40^\circ$  from acetone, then distil it under vacuum. [Beilstein 2 IV 2204.]

**Oxalic acid (2 $\text{H}_2\text{O}$ )** [6153-56-6] **M 90.0, m 101.5°, [anhydrous 144-62-7] m 189.5°,  $pK_1^{25}$  1.08 (1.37),  $pK_2^{25}$  3.55 (3.80)**. Crystallise oxalic acid from distilled water. Dry it in a vacuum over  $\text{H}_2\text{SO}_4$ . The anhydrous acid can be obtained by drying at  $100^\circ$  overnight. [Beilstein 2 IV 1819.]

**Oxaloacetic acid** [328-42-7] **M 132.1, m 160°(decarboxylates),  $pK_1^{25}$  2.22,  $pK_2^{25}$  3.89,  $pK_3^{25}$  13.0**. Crystallise it from boiling EtOAc, or from hot  $\text{Me}_2\text{CO}$ /hot  $\text{C}_6\text{H}_6$ . [Beilstein 3 IV 1808.]

**2-Oxoglutaric acid (2-oxopentane-1,5-dioic,  $\alpha$ -ketoglutaric acid)** [328-50-7] **M 146.1, m 114°, 115-117°, ( $pK_{\text{Est}}$  see oxaloacetic acid above)**. Crystallise the keto-acid repeatedly from  $\text{Me}_2\text{CO}$ /\*benzene, EtOAc or ethyl propionate. Dry it *in vacuo*. [Beilstein 3 IV 1813.]

**Oxamide** [471-46-5] **M 88.1, m >320°(dec)**. Crystallise oxamide from water, grind it and dry it in an oven at  $150^\circ$ . [Beilstein 2 IV 1860.]

**Palmitic acid anhydride (hexadecanoic anhydride)** [623-65-4] **M 494.9, m 63-64°, 64°,  $d_4^{82}$  0.838,  $n_D^{68}$  1.436**. It is moisture sensitive and hydrolyses in water. Purify it by refluxing with acetic anhydride for 1 hour, evaporating and freeing the residue of acetic acid and anhydride by drying the residue at high vacuum and recrystallising from petroleum ether at low temperature. [Beilstein 2 IV 1181.]

**Paraffin (oil)** [8012-95-1]  **$d_4^{20}$  0.880,  $n_D^{20}$  1.482**. Treat the oil with fuming  $\text{H}_2\text{SO}_4$  (care), then wash it with water and dilute aqueous  $\text{NaOH}$ , then percolate it through activated silica gel.

**Paraffin Wax.** Melt the wax in the presence of NaOH, wash it with water until all of the base had been removed. The paraffin is allowed to solidify after each wash. Finally, 5g of paraffin is melted by heating it on a water-bath, then shaken for 20-30 minutes with 100ml of boiling water and and dry the melt under vacuum.

**Pelargonic acid (nonanoic acid)** [112-05-0] M 158, m 15°, b 98.9°/1mm, 225°/760mm, pK<sup>25</sup> 4.96. Esterify the acid with ethylene glycol and distil the ester. (This removes dibasic acids as undistillable residues.) The acid is regenerated by hydrolysing the ester in the usual way and is distilled *in vacuo*. [Beilstein 2 IV 1018.]

**Pelargonitrile (octyl cyanide)** [2243-27-8] M 139.2, m -34°, b 92°/10mm, 224°, d<sub>4</sub><sup>20</sup> 0.818, n<sub>D</sub><sup>20</sup> 1.4255. Stir the nitrile with P<sub>2</sub>O<sub>5</sub> (~5%), distil it from P<sub>2</sub>O<sub>5</sub> and redistil it under a vacuum. IR should have CN but no OH bands. [Beilstein 2 IV 1204.]

**Pelargonyl chloride (nonanoyl chloride)** [764-85-2] M 176.7, b 88°/12mm, d<sub>4</sub><sup>20</sup> 0.941, n<sub>D</sub><sup>20</sup> 1.436. Reflux it with acetyl chloride (~3 volumes) for 1 hour, then distil off AcCl followed by the nonanoyl chloride at ~12mm. It is moisture sensitive and should be stored in sealed ampoules. [Beilstein 2 IV 1023.]

**Pentabromoacetone** [79-49-2] M 452.6, m 76°, pK<sup>25</sup> 8.0 (MeOH), pK<sub>Est</sub> ~4.6 (H<sub>2</sub>O). Crystallise it from Et<sub>2</sub>O, EtOH or aqueous EtOH (m 73.2°) and sublime it. Its solubility in H<sub>2</sub>O is 0.01mg/100ml. [Beilstein 1 H 659, 1 I 345, 1 III 2753, 1 IV 3226.]

**Pentachloroethane (pentalin)** [76-01-7] M 202.3, b 69°/37mm, 152.2°/64mm, 162.0°/~760mm, d<sub>4</sub><sup>20</sup> 1.678, n<sup>15</sup> 1.50542. Usual impurities include trichloroethylene. It partially decomposes if it is distilled at atmospheric pressure. Drying it with CaO, KOH or sodium is unsatisfactory because of the elimination of the elements of HCl. It can be purified by steam distillation, or by washing with conc H<sub>2</sub>SO<sub>4</sub>, water, and then aqueous K<sub>2</sub>CO<sub>3</sub>, drying with solid K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, and fractionally distilling under reduced pressure. [Beilstein 2 IV 147.]

**Pentadecafluoro octanoic acid (perfluorocaprylic acid)** [335-67-1] M 414.1, m 54.9-55.6°, b 189°/736mm, pK<sub>Est</sub> <0. Crystallise the acid from CCl<sub>4</sub> and toluene, and distil it. It forms micelles in H<sub>2</sub>O and the solubility is 1% in H<sub>2</sub>O. The *acid chloride* has b 129-130°/744mm. The *amide* has m 138°. [Bernett & Zisman *J Phys Chem* 63 1911 1959, Bro & Sperati *J Polym Sci* 38 289 1959, Beilstein 2 IV 994.]

**Pentadecanoic acid** [1002-84-2] M 242.4, m 51-53°, 80°, b 158°/1mm, 257°/760mm, d<sub>4</sub><sup>80</sup> 0.8424, pK<sub>Est</sub> ~5.0. Crystallise the acid from Et<sub>2</sub>O and distil it *in vacuo*. It is very hygroscopic. See the purification of palmitic acid. [Beilstein 2 IV 1147.]

**Pentadecanolide (1-oxacyclohexadecan-2-one, pentadecanoic- $\omega$ -lactone, 15-hydroxypentadecanoic lactone, exaltolide, Tibetolide)** [106-02-5] M 240.4, m 34-36°, 37-37.5°, 37-38°, b 102-103°/0.03mm, 112-114°/0.2mm, 137°/2mm, 169°/10-11mm, d<sub>4</sub><sup>40</sup> 0.9401. It has been recrystallised from MeOH (4 parts) at -15°, and distilled under high vacuum. [Hundiecker & Erlbach *Chem Ber* 80 135 1947, Galli & Mandolini *Org Synth* 58 100 1978, Demole & Enggist *Helv Chim Acta* 11 2318 1978, Beilstein 17/9 V 106.]

**Penta-1,3-diene** [*cis*: 1574-41-0, *trans*: 2004-70-8] M 68.1, b 42°, d<sub>4</sub><sup>20</sup> 0.680, n<sub>D</sub><sup>20</sup> 1.4316. Distil the diene from NaBH<sub>4</sub>. Purify it also by preparative gas chromatography. [Reimann et al. *J Am Chem Soc* 108 5527 1986, Beilstein 1 IV 994.]

**Penta-1,4-diene** [591-93-5] M 68.1, b 25.8-26.2°/756mm, d<sub>4</sub><sup>20</sup> 0.645, n<sub>D</sub><sup>20</sup> 1.3890. Distil it from NaBH<sub>4</sub>. Purify it by preparative gas chromatography or distillation and stabilise it with 0.1% of 2,6-di-*tert*-butyl-*p*-cresol. [Reimann et al. *J Am Chem Soc* 108 5527 1986, Beilstein 1 IV 998.]

**Pentaethylenehexamine** [4067-16-7] M 232.4, d<sub>4</sub><sup>20</sup> 0.950, n<sub>D</sub><sup>20</sup> 1.510, pK<sub>1</sub> 1.2, pK<sub>2</sub> 2.7, pK<sub>3</sub> 4.3, pK<sub>4</sub> 7.8, pK<sub>5</sub> 9.1, pK<sub>6</sub> 9.9 (all estimated). Fractionally distil it twice at 10-20mm, the fraction boiling at 220-250° being collected. It can be further purified *via* the *hydrochloride*. Its solution in MeOH (40ml of base in 250ml) is cooled in an ice-bath and conc HCl (~50ml) is added dropwise with stirring. The precipitated *hydrochloride* is

filtered off, washed with Me<sub>2</sub>CO, and Et<sub>2</sub>O, then dried in a vacuum desiccator. The free base is then obtained by basification, extraction into Et<sub>2</sub>O, drying (NaOH), filtering, evaporating and distilling the residue as before. It forms a Cu complex [Cu(C<sub>10</sub>H<sub>28</sub>N<sub>6</sub>)<sub>2</sub>]<sup>2+</sup>. [Jonassen et al. *J Am Chem Soc* **79** 4279 1957, *Beilstein* **4** IV 1245.]

**2,2,3,3,3-Pentafluoropropan-1-ol** [422-05-9] **M 150.1, b 80°/~760mm, d<sub>4</sub><sup>20</sup> 1.507, n<sub>D</sub><sup>20</sup> 1.288, pK<sup>25</sup> 12.74.** Shake the alcohol with alumina for 24 hours, dry with anhydrous K<sub>2</sub>CO<sub>3</sub>, and distil it, collect the middle fraction (**b 80-81°**) and redistil it. [*Beilstein* **1** IV 1438.]

**n-Pentane** [109-66-0] **M 72.2, b 36.1°, d<sub>4</sub><sup>20</sup> 0.626, n<sub>D</sub><sup>25</sup> 1.35472.** Stir the pentane with successive portions of conc H<sub>2</sub>SO<sub>4</sub> until there is no further coloration during 12 hours, then with 0.5N KMnO<sub>4</sub> in 3M H<sub>2</sub>SO<sub>4</sub> for 12 hours, wash with water and aqueous NaHCO<sub>3</sub>. Dry it with MgSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, then P<sub>2</sub>O<sub>5</sub> and fractionally distil it through a column packed with glass helices. It is also purified by passage through a column of silica gel, followed by distillation and storage with sodium hydride. An alternative purification is by azeotropic distillation with MeOH, which is subsequently washed out from the distillate (using water), followed by drying and re-distilling. For removal of carbonyl-containing impurities, see *n-heptane*. Also purify it by fractional freezing (*ca* 40%) on a copper coil through which cold air is passed, then wash with conc H<sub>2</sub>SO<sub>4</sub> and fractionally distil it. [*Beilstein* **1** IV 303.]

**Pentane-1-thiol** [110-66-7] **M 104.2, m -76°, b 122.9°/697.5mm, d<sup>25</sup> 0.8375, pK<sub>Est</sub> ~10.1.** Dissolve the thiol in aqueous 20% NaOH, then extract with a small amount of diethyl ether. The aqueous solution is acidified slightly with 15% H<sub>2</sub>SO<sub>4</sub>, and the thiol is distilled out, dried with CaSO<sub>4</sub> or CaCl<sub>2</sub>, and fractionally distilled under nitrogen. [Ellis & Reid *J Am Chem Soc* **54** 1674 1932, *Beilstein* **1** IV 1453.]

**(±)-Pentan-2-ol** [6032-29-7] **M 88.2, b 119.9°/~760mm, d<sub>4</sub><sup>20</sup> 0.810, n<sub>D</sub><sup>20</sup> 1.41787, n<sub>D</sub><sup>25</sup> 1.4052.** Reflux it with CaO, distil it, then reflux it with magnesium and again fractionally distil it. [*Beilstein* **1** IV 1655.]

**Pentan-3-ol** [584-02-1] **M 88.2, b 116.2°, d<sub>4</sub><sup>20</sup> 0.819, n<sub>D</sub><sup>25</sup> 1.4072.** Reflux the alcohol with CaO, distil, then reflux it with magnesium and again fractionally distil it. [*Beilstein* **1** IV 1662.]

**Pentan-3-one** see **diethyl ketone** above.

**Pent-2-ene (mixed isomers)** [109-68-2] **M 70.1, b 36.4°, d<sub>4</sub><sup>20</sup> 0.650, n<sub>D</sub><sup>20</sup> 1.38003, n<sub>D</sub><sup>25</sup> 1.3839.** Reflux the mixture with sodium wire, then fractionally distil it twice through a Fenske (glass helices packing) column. [*Beilstein* **1** IV 815.]

**cis-Pent-2-ene** [627-20-3] **M 70.1, b 37.1°, d<sub>4</sub><sup>20</sup> 0.657, n<sub>D</sub><sup>25</sup> 1.3798.** Dry it with sodium wire and fractionally distil it, or purify it by azeotropic distillation with MeOH, followed by washing out the MeOH with water, drying and distilling. Also purify it by chromatography through silica gel and alumina [Klassen & Ross *J Phys Chem* **91** 3668 1987]. [*Beilstein* **1** IV 814.]

**trans-Pent-2-ene** [646-04-8] **M 70.1, b 36.5°, d<sub>4</sub><sup>20</sup> 0.6482, n<sub>D</sub><sup>20</sup> 1.3793.** It is treated as above and washed with water, dried over anhydrous Na<sub>2</sub>CO<sub>3</sub>, and fractionally distilled. The middle cut is purified by two passes of fractional melting. [*Beilstein* **1** IV 814.]

**Pent-2-yne** [627-21-4] **M 68.1, b 26°/2.4mm, 56.1°/760mm, d<sub>4</sub><sup>20</sup> 0.710, n<sub>D</sub><sup>25</sup> 1.4005.** It is stood with, then distilled at low pressure from sodium or NaBH<sub>4</sub>. [*Beilstein* **1** III 958, **1** IV 992.]

**Perfluorobutyric acid (heptafluorobutyric acid)** [375-22-4] **M 214.0, m -17.5°, b 120°/735mm, d<sub>4</sub><sup>20</sup> 1.651, n<sub>D</sub><sup>16</sup> 1.295, pK<sup>25</sup> -0.17.** Fractionally distil the acid twice in an Oldershaw column with an automatic vapour-dividing head, the first distillation being in the presence of conc H<sub>2</sub>SO<sub>4</sub> as a drying agent. (Take care with the hot acid.) [*Beilstein* **2** IV 810.]

**Perfluoroheptane (hexadecafluoroheptane)** [335-57-9] **M 388.1, b 99-101°, d<sub>4</sub><sup>25</sup> 1.7200.** Purify it as for *perfluorodimethylhexane*. Other procedures include shaking with H<sub>2</sub>SO<sub>4</sub>, washing with water, and drying with

$P_2O_5$  for 48 hours then fractionally distilling. *Alternatively*, it has been refluxed for 24 hours with saturated acid  $KMnO_4$  (to oxidise and remove hydrocarbons), then neutralised, steam distilled, dried with  $P_2O_5$ , and passed slowly through a column of dry silica gel. It has been purified by fractional crystallisation using partial freezing. [Beilstein 1 IV 388.]

**Perfluoro-*n*-hexane (tetradecafluorohexane)** [355-42-0] M 338.1, m  $-4^\circ$ , b 58-60 $^\circ$ ,  $d_4^{20}$  1.684. Purify the fluorohexane by fractional freezing. The methods described for *perfluoroheptane* should be applicable here. [Beilstein 1 IV 348.]

**Perfluorononane (eicosafluorononane)** [375-96-2] M 488.1, b 126-127 $^\circ$ ,  $d_4^{20}$  1.80,  $n_D^{20}$  1.275. Purify as for *perfluorodimethylcyclohexane*. [Beilstein 1 III 505.]

**Perfluoropropyl iodide (heptafluoro-1-iodopropane)** [754-34-7] M 295.9, b 41 $^\circ$ ,  $d_4^{20}$  2.13,  $n_D^{20}$  1.339. Purify the fluoro-iodide by fractional distillation. Store it over Cu as stabiliser. [Beilstein 1 IV 225.]

**Perfluorotributylamine (heptacosafuorotributylamine)** [311-89-7] M 671.1, b 177.6 $^\circ$ /760mm,  $d_4^{20}$  1.881,  $n_D^{20}$  1.291,  $pK_{Est}$  ~5.0. Purify it as for *perfluorodimethylcyclohexane* (see [335-27-3]); see also *perfluorotripropylamine* below [Haszeldine *J Chem Soc* 102 1951]. [Beilstein 2 IV 819.]

**Perfluorotripropylamine (heneicosafuorotripropylamine)** [338-83-0] M 521.1, b 130 $^\circ$ /atm, 129.5-130.5 $^\circ$ /atm,  $d_4^{20}$  1.822,  $n_D^{20}$  1.279,  $pK_{Est}$  ~5.6. Purify it as for *perfluorodimethylcyclohexane* (see [335-27-3]) [Haszeldine *J Chem Soc* 102 1951, for azeotropes see Simons & Linevsky *J Am Chem Soc* 74 4750 1972.]  
**IRRITANT.**

**Petroleum ether** [8032-32-4] b 35-60 $^\circ$ ,  $d_4^{20}$  0.640,  $n_D^{20}$  1.363. Shake it several times with conc  $H_2SO_4$ , then 10%  $H_2SO_4$  and concentrated  $KMnO_4$  (to remove unsaturated, including aromatic, hydrocarbons) until the permanganate colour persists. Wash it with water, aqueous  $Na_2CO_3$  and again with water. Dry it with  $CaCl_2$  or  $Na_2SO_4$ , and distil it. It can be dried further using  $CaH_2$  or sodium wire. Passage through a column of activated alumina, or treatment with  $CaH_2$  or sodium, removes peroxides. For the elimination of carbonyl-containing impurities without using permanganate, see *n-heptane*. These procedures could be used for all fractions of petroleum ethers. See *skellysolve* below, p. 201.

**Rapid purification:** Pass it through an alumina column and fractionally distilling, collecting the desired boiling fraction.

**Phorone (2,6-dimethylhepta-2,5-dien-4-one)** [504-20-1] M 138.2, m 28 $^\circ$ , b 197 $^\circ$ /743mm. Crystallise phorone repeatedly from EtOH. [Beilstein 1 IV 3564.]

**Pimelic acid (heptane-1,7-dioic acid)** [111-16-0] M 160.2, m 105-106 $^\circ$ ,  $pK_1^{25}$  4.46,  $pK_2^{25}$  5.58. Crystallise the acid from water or from \*benzene containing 5% diethyl ether. [Beilstein 1 IV 2003.]

**Pinacol (hexahydrate)** [6091-58-3 (6 $H_2O$ ), 76-09-5 (anhydrous)] M 194.3, m 46.5 $^\circ$ , b 59 $^\circ$ /4mm. Distil pinacol, then crystallise it repeatedly from water. (See also below.) [Beilstein 1 IV 2575.]

**Pinacol (anhydrous)** [76-09-5] M 118.1, m 41.1 $^\circ$ , b 172 $^\circ$ . The hydrate is rendered anhydrous by azeotropic distillation of water with \*benzene. Recrystallise it from \*benzene or toluene/petroleum ether, absolute EtOH or dry diethyl ether. It recrystallises from water to give the *hexahydrate*. [Beilstein 1 IV 2575.]

**Pinacolone** see *tert-butyl methyl ketone*.

**Pinacolone oxime** [2475-93-6] M 115.2, m 75.5-76 $^\circ$ , 78 $^\circ$ , 78.5-79.5 $^\circ$ , 171.6 $^\circ$ /748mm. Crystallise the oxime from aqueous EtOH, EtOH (small needles) or petroleum ether (plates). [Markownikoff *Chem Ber* 32 1448 1899, Smith & Atkins *J Am Chem Soc* 60 660 1938, Whitmore et al. *J Am Chem Soc* 61 684 1939, Beilstein 1 H 694, 1 II 750, 1 III 2842, 1 IV 3310.]

**Pivalic acid (trimethylacetic acid)** [75-98-9] **M 102.1, m 35.4°, b 71-73°/0.1mm, pK<sup>25</sup> 5.03.** Fractionally distil the acid under reduced pressure, then fractionally crystallise it from its melt. Recrystallise it from \*benzene. [Beilstein 2 IV 908.]

**Pivaloyl chloride (trimethylacetyl chloride)** [3282-30-2] **M 120.6, b 57.6°/150mm, 70.5-71/250mm, 104°/754mm, 104-105°/atm, 105-108°/atm, d<sub>4</sub><sup>20</sup> 1.003, n<sub>D</sub><sup>20</sup> 1.4142.** First check the IR to see if OH bands are present. If absent, or present in small amounts, then redistil it under a moderate vacuum. If present in large amounts then treat it with oxalyl chloride or thionyl chloride and reflux for 2-3 hours, evaporate and distil the residue. **Strongly LACHRYMATORY – work in a fume cupboard.** Store it in sealed ampoules under N<sub>2</sub>. [Traynham & Battiste *J Org Chem* 22 1551 1957, Grignard reactions: Whitmore et al. *J Am Chem Soc* 63 647 1941, Beilstein 2 IV 912.]

**Polyacrylonitrile** [25014-41-9]. Precipitate it from dimethylformamide by addition of MeOH.

**Poly(diallyldimethylammonium) chloride** [26062-79-3]. Precipitate it from water with acetone, and dry the salt in a vacuum for 24 hours. [Hardy & Shriner *J Am Chem Soc* 107 3822 1985.]

**Polyethylene** [9002-88-4]. Crystallise it from thiophen-free \*benzene and dry it over P<sub>2</sub>O<sub>5</sub> under vacuum.

**Polymethyl acrylate** [9003-21-8]. Precipitate it from a 2% solution in acetone by addition of water.

**Polyvinyl acetate** [9003-20-7]. Precipitate it from acetone by addition of *n*-hexane.

**Polyvinyl chloride** [9002-81-2]. Precipitate it from cyclohexanone by addition of MeOH.

**Propane** [74-98-6] **M 44.1, m -189.7, b -42.1°/760mm, d<sub>4</sub><sup>20</sup> 0.5005, n<sub>D</sub><sup>20</sup> 1.2898.** Purify propane by bromination of the olefinic contaminants. Propane is treated with bromine for 30 minutes at 0°. Unreacted bromine is quenched, and the propane is distilled through two -78° traps and collected at -196° [Skell et al. *J Am Chem Soc* 108 6300 1986]. It autoignites at 450° and the flash point is -104°. It is highly **FLAMMABLE** and is available in metal cylinders. [Beilstein 1 H 103, 1 I 33, 1 II 71, 1 III 204, 1 IV 175.]

**Propane-1,2-diamine (propylenediamine)** [78-90-0] **M 74.1, b 120.5°/~760mm, d<sub>4</sub><sup>20</sup> 0.868, n<sub>D</sub><sup>20</sup> 1.446, pK<sub>1</sub><sup>25</sup> 6.61, pK<sub>2</sub><sup>25</sup> 9.82.** Purify the diamine by azeotropic distillation with toluene. Then distil it. Store it in a CO<sub>2</sub> free atmosphere. [Horton et al. *Anal Chem* 27 269 1955, Beilstein 4 IV 1255.]

**(±)-Propane-1,2-diol (propyleneglycol)** [57-55-6] **M 76.1, m -60°, b 45.5°/1.0mm, 70.8°/5.0mm, 83.2°/10mm, 96.4°/20mm, 104°/32mm, 111.2°/40mm, 119.9°/60mm, 132.0°/100mm, 149.7°/200mm, 168.1°/400mm, 188.2°/atm, d<sub>4</sub><sup>20</sup> 1.040, n<sub>D</sub><sup>20</sup> 1.433.** Dry the diol over Na<sub>2</sub>SO<sub>4</sub>, decant and distil it under reduced pressure. It is soluble in H<sub>2</sub>O, Me<sub>2</sub>CO, CHCl<sub>3</sub>, many oils, is used as an antifreeze in dairies and breweries; and is a substitute for glycerol or ethylene glycol. [Beilstein 1 IV 2468.]

***R*(-)- and *S*(+)- Propane-1,2-diol (*R*- and *S*- propylene glycol)** [*R*(-) 4254-14-2; *S*(+) 4254-15-3] **M 76.1, b 78.5°/10mm, 94-96°/14mm, 186-188°/765mm, d<sub>4</sub><sup>20</sup> 1.036, n<sub>D</sub><sup>20</sup> 1.432, [α]<sub>D</sub><sup>25</sup> (-) or (+) 17.5° (neat).** The laevo *R*-enantiomer can be obtained from reduction of hydroxyacetone by yeast [Levene & Walti *Org Synth Coll Vol II* 545 1943], and the dextro *S*-enantiomer is obtained by reduction of *S*(+)-lactic acid. Thus (+)-lactic acid (33.4g, 370mmol) in dry THF (200ml) is added dropwise to a suspension of LAH (32.1g, 850mmol) in dry THF (500ml) at 0° during 2 hours, allowed to warm to 25°, refluxed for 2 hours, cooled, then quenched by careful addition of ice-cold H<sub>2</sub>O (32ml), followed by 4N NaOH (32ml) and again H<sub>2</sub>O (96ml). The white precipitate is coagulated by refluxing the mixture for 0.5 hours, filtered, washed with THF, the white cake is slurred with hot THF (4 x ~100ml), the combined THF layers are evaporated *in vacuo* (~40°), and the residual oil is distilled through a vacuum jacketed Vigreux column under vacuum to give the *S*-diol as a colourless liquid (23.9g, 85%). [Fryzuk & Bosnich *J Am Chem Soc* 100 5491 1978.]

**Propane-1,3-diol** [504-63-2] **M 76.1, b 110-122°/12mm, d<sub>4</sub><sup>20</sup> 1.053, n<sub>D</sub><sup>18.5</sup> 1.4398.** Dry this diol with K<sub>2</sub>CO<sub>3</sub>

and distil it under reduced pressure. More extensive purification involves conversion with benzaldehyde to 2-phenyl-1,3-dioxane (**m** 47-48°) which is subsequently decomposed by shaking with 0.5M HCl (3ml/g) for 15 minutes and standing overnight at room temperature. After neutralisation with K<sub>2</sub>CO<sub>3</sub>, the benzaldehyde is removed by steam distillation and the diol is recovered from the remaining aqueous solution by continuous extraction with CHCl<sub>3</sub> for 1 day. The extract is dried with K<sub>2</sub>CO<sub>3</sub>, the CHCl<sub>3</sub> is evaporated and the diol is distilled. [Foster et al. *Tetrahedron* **6** 177 1961, *Beilstein* **1** IV 2493.]

**S-(-)-Propane-1,2-diol bis(*p*-toluenesulfonate)** [60434-71-1] **M** 384.5, **m** 68-70° (62°), [α]<sub>D</sub><sup>20</sup> -20° (**c** 1, CHCl<sub>3</sub>), is prepared in much the same way as for *S,S*-butane-2,3-diol bis-tosylate (see above) from *S*-(+)-propane-1,2-diol (310mmol) in dry pyridine (30ml) and toluene-*p*-sulfonyl chloride (700mmol, recrystallised from hexane) in dry pyridine (135ml) at 0° during 0.5 hours, then at 25° for 17 hours. The crude ester is purified by dissolving in CH<sub>2</sub>Cl<sub>2</sub> and cyclohexane is added to cloud point at 40°, then allowed to crystallise at 25° when more cyclohexane is added and set aside (5°/12 hours) to give the *bis-tosylate* as white feathery crystals which are dried over CaCl<sub>2</sub> *in vacuo* (114g, 95%). [Fryzuk & Bosnich *J Am Chem Soc* **100** 5491 1978.] This *S*-(–)-2,3-ditosylate provides the chiral ligand *R*-(+)- [cf 15629-92-2], which involves inversion of configuration at the chiral centre.

**Propane-1-thiol** [107-03-9] **M** 76.1, **b** 65.3°/702mm, **d**<sub>4</sub><sup>25</sup> 0.83598, **n**<sub>D</sub><sup>25</sup> 1.43511, **pK**<sup>20</sup> 10.82. Purify the thiol by dissolving it in aqueous 20% NaOH, extracting with a small amount of \*benzene and steam distilling until clear. After cooling, the solution is acidified slightly with 15% H<sub>2</sub>SO<sub>4</sub>, and the thiol is distilled out, dried with anhydrous CaSO<sub>4</sub> or CaCl<sub>2</sub>, and fractionally distilled under nitrogen. [Mathias & Filho *J Phys Chem* **62** 1427 1958.] Also purify it by liberating the mercaptan by adding dilute HCl to the residue remaining after steam distilling. After direct distillation from the flask, and separation of the water, the mercaptan is dried (Na<sub>2</sub>SO<sub>4</sub>) and distilled under nitrogen. [*Beilstein* **1** IV 1449.]

**Propane-2-thiol (Isopropyl mercaptan)** [75-33-2] **M** 76.1, **b** 49.8°/696mm, **d**<sub>4</sub><sup>25</sup> 0.80895, **n**<sub>D</sub><sup>25</sup> 1.42154, **pK**<sup>25</sup> 10.86. Purify it as for propane-1-thiol above. [*Beilstein* **1** IV 1498.]

**Propargyl alcohol (2-propyn-1-ol)** [107-19-7] **M** 56.1, **b** 54°/57mm, 113.6°/760mm, **d**<sub>4</sub><sup>20</sup> 0.947, **n**<sub>D</sub><sup>20</sup> 1.432. The commercial material contains a stabiliser. An aqueous solution of propargyl alcohol can be concentrated by azeotropic distillation with butanol or butyl acetate. Dry it with K<sub>2</sub>CO<sub>3</sub> and distil it under reduced pressure, in the presence of about 1% succinic acid, through a glass helices-packed column. [*Beilstein* **1** IV 2214.]

**Propargyl chloride (3-chloropropyne)** [624-65-7] **M** 74.5, **b** 58°/760mm, 65°/760mm, **d**<sub>4</sub><sup>20</sup> 1.03, **n**<sub>D</sub><sup>20</sup> 1.435. Purify the chloride by fractional distillation at atmospheric pressure. Note that a possible impurity is propargyl alcohol which has **b** 114-115°/760mm (see above). [Henry *Chem Ber* **8** 398 1875.] **HIGHLY TOXIC** and **FLAMMABLE**. [*Beilstein* **1** IV 960.]

**Propene (propylene)** [115-07-1] **M** 42.1, **m** -185.2°, **b** -47.8°/750mm, **d**<sub>4</sub><sup>20</sup> 0.519, **n**<sup>-71</sup> 1.357. Purify it by freeze-pump-thaw cycles and trap-to-trap distillation. [*Beilstein* **1** IV 725.]

**β-Propiolactone** see oxetan-2-one in “Heterocyclic Compounds”, in this Chapter.

**Propionaldehyde** [123-38-6] **M** 58.1, **b** 48.5-48.7°, **d**<sub>4</sub><sup>20</sup> 0.804, **n**<sub>D</sub><sup>20</sup> 1.3733, **n**<sub>D</sub><sup>25</sup> 1.37115. Dry the aldehyde with CaSO<sub>4</sub> or CaCl<sub>2</sub>, and fractionally distil it under nitrogen or in the presence of a trace of hydroquinone (to retard oxidation). Blacet and Pitts [*J Am Chem Soc* **74** 3382 1952] repeatedly distilled the middle fraction in a vacuum until it no longer gave a solid polymer when cooled to -80°. It is stored with CaSO<sub>4</sub>. [*Beilstein* **1** IV 3165.]

**Propionamide** [79-05-0] **M** 73.1, **m** 79.8-80.8°, **pK**<sup>24</sup> -0.9 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise it from acetone, \*benzene, CHCl<sub>3</sub>, water or acetone/water, then dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or conc H<sub>2</sub>SO<sub>4</sub>. [*Beilstein* **2** H 243, **2** I 108, **2** II 223, **2** III 542, **2** IV 725.]

**Propionic acid** [79-09-4] M 74.1, b 141°, d<sub>4</sub><sup>20</sup> 0.992, n<sub>D</sub><sup>20</sup> 1.3865, n<sub>D</sub><sup>25</sup> 1.3843, pK<sub>1</sub><sup>25</sup> –6.8 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>25</sup> 4.88. Dry the acid with Na<sub>2</sub>SO<sub>4</sub> or by fractional distillation, then redistil after refluxing with a few crystals of KMnO<sub>4</sub>. An alternative purification uses conversion to the ethyl ester, fractional distillation and hydrolysis. [Bradbury *J Am Chem Soc* 74 2709 1952.] Propionic acid can also be heated for 0.5 hour with an amount of benzoic anhydride equivalent to the amount of water present (in the presence of CrO<sub>3</sub> as catalyst), followed by fractional distillation. [Cham & Israel *J Chem Soc* 196 1960, *Beilstein* 2 IV 695.]

**Propionic anhydride** [123-62-6] M 130.2, b 67°/18mm, 168°/780mm, d<sub>4</sub><sup>20</sup> 1.407, n<sub>D</sub><sup>20</sup> 1.012. Shake the anhydride with P<sub>2</sub>O<sub>5</sub> for several minutes, then distil. [*Beilstein* 2 IV 722.]

**Propionitrile** [107-12-0] M 55.1, b 97.2°, d<sub>4</sub><sup>20</sup> 1.407, n<sub>D</sub><sup>15</sup> 1.36812, n<sub>D</sub><sup>30</sup> 1.36132. Shake the nitrile with dilute HCl (20%), or with concentrated HCl until the odour of isonitrile has gone, then wash with water, and aqueous K<sub>2</sub>CO<sub>3</sub>. After a preliminary drying with silica gel or Linde type 4A molecular sieves, it is stirred with CaH<sub>2</sub> until hydrogen evolution ceases, then decant and distil from P<sub>2</sub>O<sub>5</sub> (not more than 5g/L, to minimise gel formation). Finally, it is refluxed with, and slowly distilled from CaH<sub>2</sub> (5g/L), taking precautions to exclude moisture. [*Beilstein* 2 IV 728.]

***n*-Propyl acetate** [109-60-4] M 102.1, b 101.5°, d<sub>4</sub><sup>20</sup> 0.887, n<sub>D</sub><sup>20</sup> 1.38442, pK<sub>25</sub> –7.18 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>). Wash the ester with saturated aqueous NaHCO<sub>3</sub> until neutral, then with saturated aqueous NaCl. Dry it with MgSO<sub>4</sub> and fractionally distil it. [*Beilstein* 2 IV 138.]

***n*-Propyl alcohol (1-propanol)** [71-23-8] M 60.1, b 97.2°, d<sub>4</sub><sup>25</sup> 0.79995, n<sub>D</sub><sup>20</sup> 1.385, pK<sub>25</sub> 16.1. The main impurities in *n*-propyl alcohol are usually water and 2-propen-1-ol, reflecting the commercial production by hydration of propene. Water can be removed by azeotropic distillation either directly (azeotrope contains 28% water) or by using a ternary system, e.g. by also adding \*benzene. *Alternatively*, for removal of gross amounts of water, reflux over CaO for several hours is desirable, followed by distillation and a further drying. To obtain more nearly anhydrous alcohol, suitable drying agents are firstly NaOH, CaSO<sub>4</sub> or K<sub>2</sub>CO<sub>3</sub>, then CaH<sub>2</sub>, aluminium amalgam, magnesium activated with iodine, or a small amount of sodium. *Alternatively*, the alcohol can be refluxed with *n*-propylsuccinate or phthalate in a method similar to the one described under EtOH. Allyl alcohol is removed by adding bromine (15ml/L), and then fractionally distilling from a small amount of K<sub>2</sub>CO<sub>3</sub>. Propionaldehyde, also formed in the bromination, is removed as the 2,4-dinitrophenylhydrazone. *n*-Propyl alcohol can be dried down to 20 ppm of water by passage through a column of pre-dried molecular sieves (type 3 or 4A, heated for 3 hours at 300°) in a current of nitrogen. Distillation from sulfanilic or tartaric acids removes impurities.

Albrecht [*J Am Chem Soc* 82 3813 1960] obtained spectroscopically pure material by heating with charcoal to 50-60°, filtering and adding 2,4-dinitrophenylhydrazine and a few drops of conc H<sub>2</sub>SO<sub>4</sub>. After standing for several hours, the mixture is cooled to 0°, filtered and distilled in a vacuum. Gold and Satchell [*J Chem Soc* 1938 1963] heated *n*-propyl alcohol with 3-nitrophthalic anhydride at 76-110° for 15 hours, then recrystallised the resulting ester from H<sub>2</sub>O, \*benzene/petroleum ether (b 100-120°)(3:1), and \*benzene. The ester was hydrolysed under reflux with aqueous 7.5M NaOH for 45 minutes under nitrogen, followed by distillation (also under nitrogen). The fraction with b 87-92° is dried with K<sub>2</sub>CO<sub>3</sub> and stirred under reduced pressure in the dark over 2,4-dinitrophenylhydrazine, then freshly distilled. Also purify it by adding 2g NaBH<sub>4</sub> to 1.5L of alcohol, gently flushing with argon and refluxing for 1 day at 50°. Then 2g of freshly cut sodium (washed with propanol) is added and refluxed for one day, and finally distilled, taking the middle fraction [Jou & Freeman *J Phys Chem* 81 909 1977]. [*Beilstein* 1 IV 1413.]

***n*-Propylamine** [107-10-8] M 59.1, b 48.5°, d<sub>4</sub><sup>20</sup> 0.716, n<sub>D</sub><sup>20</sup> 1.38815, pK<sub>25</sub> 10.69. Distil the amine from zinc dust, under reduced pressure, in an atmosphere of nitrogen. [*Beilstein* 4 IV 464.]

***n*-Propyl bromide.** [106-94-5] M 123.0, b 71.0°, d<sub>4</sub><sup>20</sup> 1.354, n<sub>D</sub><sup>15</sup> 1.43695, n<sub>D</sub><sup>25</sup> 1.43123. Likely contaminants include *n*-propyl alcohol and isopropyl bromide. The simplest purification procedure uses drying with MgSO<sub>4</sub> or CaCl<sub>2</sub> (with or without a preliminary wash of the bromide with aqueous NaHCO<sub>3</sub>, then water), followed by fractional distillation away from bright light. Chien and Willard [*J Am Chem Soc* 79 4872 1957] bubbled a stream of oxygen containing 5% ozone through *n*-propyl bromide for 1 hour, then stirred it with 3% hydrogen

peroxide solution, neutralised it with aqueous  $\text{Na}_2\text{CO}_3$ , washed it with distilled water and dried it. This was followed by vigorous stirring with 95%  $\text{H}_2\text{SO}_4$  until fresh acid did not discolour within 12 hours. The propyl bromide was separated, neutralised, washed, dried with  $\text{MgSO}_4$  and fractionally distilled. The centre cut was stored in the dark. Instead of ozone, Schuler and McCauley [*J Am Chem Soc* **79** 821 1957] added bromine and stored it for 4 weeks, the bromine then being extracted with aqueous  $\text{NaHSO}_3$  before the sulfuric acid treatment was applied and finally distilled. Further purification is by preparative gas chromatography on a column packed with 30% SE-30 (General Electric ethylsilicone rubber) on 42/60 Chromosorb P at  $150^\circ$  and 40psi, using helium as carrier gas. [Chu *J Phys Chem* **41** 226 1964, *Beilstein* **1** IV 205.]

***n*-Propyl chloride** [540-54-5] **M 78.5, b 46.6°**,  $d_4^{20}$  **0.890**,  $n_D^{20}$  **1.3880**. Dry the chloride with  $\text{MgSO}_4$  and fractionally distil. It can be more extensively purified using extraction with  $\text{H}_2\text{SO}_4$  as for *n*-propyl bromide. Alternatively, Chien and Willard [*J Am Chem Soc* **75** 6160 1953] passed a stream of oxygen containing about 5% ozone through the *n*-propyl chloride for three times as long as was needed to cause the first coloration of starch iodide paper by the exit gas. After washing with aqueous  $\text{NaHCO}_3$  to hydrolyse ozonides and remove organic acids, the chloride was dried with  $\text{MgSO}_4$  and fractionally distilled. [*Beilstein* **1** IV 189.]

**Propylene glycol diacetate (1,2-diacetoxyp propane, propylene diacetate)** [623-84-7] **M 160.2, b 84-85°/12mm, 191°/760mm**,  $d_4^{25}$  **1.050**,  $n_D^{20}$  **1.414**. Wash the ester with aqueous  $\text{NaHCO}_3$  in the presence of solid  $\text{NaCl}$ , dry it with  $\text{MgSO}_4$  and fractionally distil it. [*Beilstein* **2** H 142, **2** II 156, **2** III 312, **2** IV 220.]

**Propylene glycol 1-methylether (1-methoxy-2-propanol)** [*RS*(±) 107-98-2, *R*(+) 4984-22-9, *S*(-) 26550-55-0] **M 90.1, b 119-120°/atm for *RS*, 131-132°/atm for *R* or *S***,  $d_4^{25}$  **0.922**,  $n_D^{20}$  **1.403**. Wash the ethers with aqueous  $\text{NaHCO}_3$  in the presence of solid  $\text{NaCl}$ , dry them with  $\text{MgSO}_4$  and fractionally distil them. The *RS*-acetate [108-65-6] **M 132.2** has **b 145-146°/atm**. The *R*(+) and *S*(-) enantiomers have  $[\alpha]_D^{20} \pm 20.5^\circ$  (c 10,  $\text{H}_2\text{O}$ ). [*Beilstein* **1** II 536, **1** III 2146, **1** IV 2471.]

***n*-Propyl ether (dipropyl ether)** [111-43-3] **M 102.2, b 90.1°**,  $d_4^{20}$  **0.740**,  $n_D^{15}$  **1.38296**,  $n_D^{20}$  **1.3803**,  $pK^{25}$  **-4.40** (aqueous  $\text{H}_2\text{SO}_4$ ). Purify the ether by drying with  $\text{CaSO}_4$ , by passage through an alumina column (to remove peroxides), and by fractional distillation. [*Beilstein* **1** III 2146, **1** IV 1422.]

**Propyl formate** [110-74-7] **M 88.1, b 81.3°**,  $d_4^{20}$  **0.9058**,  $n_D^{20}$  **1.3779**. Distil the formate, then wash it with saturated aqueous  $\text{NaCl}$ , and with saturated aqueous  $\text{NaHCO}_3$  in the presence of solid  $\text{NaCl}$ , dry it with  $\text{MgSO}_4$  and fractionally distil it. [*Beilstein* **2** IV 26.]

***n*-Propyl iodide (1-iodopropane)** [107-08-4] **M 170.0, b 38°/80mm, 102.5°/atm**,  $d_4^{20}$  **1.745**,  $n_D^{20}$  **1.5041**. It should be distilled first under reduced pressure to avoid decomposition. Dry the iodide with  $\text{MgSO}_4$  or silica gel and fractionally distil it. Store it under nitrogen with mercury in a brown bottle. Prior to distillation, free iodine can be removed by shaking with copper powder or by washing with aqueous  $\text{Na}_2\text{S}_2\text{O}_3$  and drying. Alternatively, the *n*-propyl iodide can be treated with bromine, then washed with aqueous  $\text{Na}_2\text{S}_2\text{O}_3$ , dried and distilled. See also *n*-butyl iodide. [*Beilstein* **1** IV 222.]

**2-Propylpentanoic acid (Valproic acid, VPA, di-*n*-propylacetic acid)** [99-66-1] **M 144.2, b 120-121°/14mm, 219.5°/760mm, 220°/atm**,  $d_4^{20}$  **0.91**,  $n_D^{25}$  **1.425**,  $pK^{25}$  **4.61**. Purify it in the same way as for propionic acid using  $\text{KMnO}_4$  (see [79-09-4]). It distils as a colourless oil with a characteristic smell and is soluble in most organic solvents. Its solubility in  $\text{H}_2\text{O}$  is 1.3mg/ml at room temperature. Its FT-IR (film) has  $\nu_{\text{max}}$  at 3572.9, 2969.9, 2885.0, 1771.8, 1465.6, 1388.2, 1102.6, 738.5, 586.5  $\text{cm}^{-1}$ ; its  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 0.92 (t,  $J = 7\text{Hz}$ , 2  $\text{CH}_3$ ), 1.40 (m, 6H), 1.60 (m, 2H), 2.29 (m, 1H), 12.2 (brs, OH); and the  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ ,  $\text{CHCl}_3$  standard) has  $\delta$  at 183.41, 45.20, 34.37, 20.59, 13.99. The *ethyl ester* [17022-31-0] **M 172.2** has **b 183°/760mm**. It forms a **1:1 Na salt (sodium valproate, Depakin among other trade names)** [1069-66-5] **M 166.2** as a white, crystalline, odourless, hygroscopic solid with the following solubilities in g/ml: 2.5 in  $\text{H}_2\text{O}$ , 0.66 in EtOH, 0.2 in MeOH, and is insoluble in organic solvents. It also forms a **2:1 Na salt [sodium hydrogen bis(2-propylpentanoate), valproate semisodium]** [76584-70-8] **M 310.4**, a **Magnesium salt (Depamag)**, and the pro-drug amide, **Valpromide (Depramide)** [2430-27-5] **M 143.2, m 123°**, **122-124°**, **125-126°**, obtained as needles by heating the acid with  $\text{NH}_3$  at  $180^\circ$ , is a bitter, crystalline white

powder [*Beilstein 2* IV 695]. These are all commercially available drugs for the treatment of epilepsy, bipolar disorders and migraine. They are **toxic** at high concentrations. [Wiemann et al. *Bull Soc Chim Fr* 189 1958, Nau et al. *Pharmacol Toxicol* 69 310 1991, Lscher *Prog Neurobiol* 58 31 1999, *Beilstein 2* H 350.]

***n*-Propyl propionate** [106-36-5] **M 120.2, b 122°, d<sub>4</sub><sup>20</sup> 0.881, n 1.393**. Treat the ester with anhydrous CuSO<sub>4</sub>, then distil it under nitrogen. [*Beilstein 2* IV 707.]

**Propyne** [74-99-7] **M 40.1, m -101.5°, b -23.2°/760mm, d<sup>-50</sup> 0.7062, n<sup>-40</sup> 1.3863**. Purify it by preparative gas chromatography. [*Beilstein 1* H 246.]

**Pyruvic acid** [127-17-3] **M 88.1, m 13°, b 65°/10mm, pK<sup>25</sup> 2.39 (2.60)**. Distil it twice, then fractionally crystallise it by partial freezing. [*Beilstein 3* IV 1505.]

**Ricinoleic acid (dl 12-hydroxyoleic acid)** [141-22-0] **M 298.5, m 7-8° (α-form), 5.0° (γ-form), n<sub>D</sub><sup>20</sup> 1.4717, pK<sub>Est</sub> ~4.5**. Purify it as the methyl acetylricinoleate [Rider *J Am Chem Soc* 53 4130 1931], fractionally distilling it at 180-185°/0.3mm, then 87g of this ester is hydrolysed by refluxing with KOH (56g), water (25ml), and MeOH (250ml) for 10 minutes. The free acid is separated on acidification and extracted into Et<sub>2</sub>O, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated and the residue is crystallised from acetone at -50°, and distilled in small batches, **b 180°/0.005mm**. The *R*-enantiomer has **m +5.5°, b 245°/10mm** and **[α]<sub>D</sub><sup>26</sup> +7.2° (c 5, Me<sub>2</sub>CO)**. [Bailey et al. *J Chem Soc* 3027 1957, *Beilstein 3* IV 1026, 1207.]

**Sebacic acid (1,10-decanedioic acid)** [111-20-6] **M 202.3, m 134.5°, pK<sub>1</sub><sup>25</sup> 4.58, pK<sub>2</sub><sup>25</sup> 5.54**. Purify sebacic acid *via* the disodium salt which, after crystallisation from boiling water (charcoal), is then converted to the free acid. The free acid is crystallised repeatedly from hot distilled water or from Me<sub>2</sub>CO/petroleum ether and dried under vacuum. [*Beilstein 2* IV 2078.]

**Sebacic acid monomethyl ester** [818-88-2] **M 216.3, m 42-43°, b 169-171°/4mm**. Recrystallise the ester from Me<sub>2</sub>CO/petroleum ether or petroleum ether at low temperature and distil it in a vacuum. [*Beilstein 2* IV 608.]

**Sebaconitrile (decanedinitrile)** [1871-96-1] **M 164.3, m 8°, b 199-200°**. Mix the nitrile with P<sub>2</sub>O<sub>5</sub> (10% by wt) distil the dinitrile from it, then redistil it. [*Beilstein 2* IV 2089.]

**Semicarbazide hydrochloride (hydrazine carboxamide hydrochloride)** [563-41-7] **M 111.5, m 173°(dec), 175°(dec), pK<sup>24</sup> 3.66**. Crystallise the salt from aqueous 75% EtOH and dry it under vacuum over CaSO<sub>4</sub>. Alternatively, crystallise it from a mixture of 3.6 mole % MeOH and 6.4 mole % of water. [Kovach et al. *J Am Chem Soc* 107 7360 1985.] It IR has  $\nu_{\max}$  at 700, 3500 cm<sup>-1</sup> [Ingersoll et al. *Org Synth Coll Vol I* 485 1941, Davison & Christie *J Chem Soc* 3389 1955, Thiele & Stange *Chem Ber* 27 33 1894, pK: Bartlett *J Am Chem Soc* 54 2853 1923]. The **free base** crystallises as prisms from absolute EtOH, **m 96°**. [Curtius & Heidenreich *Chem Ber* 27 55 1894, *Beilstein 3* IV 177.] **TOXIC ORALLY**, possible **CARCINOGEN** and **TERATOGEN**.

**Senecialdehyde (3,3-dimethylacraldehyde, 3-methyl-2-butenal)** [107-86-8] **M 84.1, b 78-80°/70mm, 133-135°/atm, d<sub>4</sub><sup>20</sup> 0.911, n<sub>D</sub><sup>20</sup> 1.428**. This flammable oil oxidises readily and should be fractionated under N<sub>2</sub>. It should be stored under N<sub>2</sub> and/or vacuum. The UV has  $\lambda_{\max}$  at 235.5nm. The *semicarbazone* has **m 221-222°** (from MeOH) and the *2,4-dinitrophenylhydrazone* has **m 184-185°** (from MeOH). [Forbes & Skilton *J Org Chem* 24 436 1959, *Beilstein 1* III 2990, 1 IV 3464.]

**Skellysolve A** is essentially *n*-pentane, **b 28-30°**,

**Skellysolve B** is essentially *n*-hexane, **b 60-68°**,

**Skellysolve C** is essentially *n*-heptane, **b 90-100°**,

**Skellysolve D** is mixed heptanes, **b 75-115°**,

**Skellysolve E** is mixed octanes, **b 100-140°**,

**Skellysolve F** is petroleum ether, **b** 30-60°,

**Skellysolve G** is petroleum ether, **b** 40-75°,

**Skellysolve H** is hexanes and heptanes, **b** 69-96°,

**Skellysolve L** is essentially octanes, **b** 95-127°. For methods of purification, see **petroleum ether**.

**Solanone** [*S(+)-trans-2-methyl-5-isopropyl-1,3-nonan-8-one*] [1937-54-8] **M 194.3, b 60°/1mm, n<sub>D</sub><sup>20</sup> 1.4755, [α]<sub>D</sub><sup>20</sup> +14° (neat)**. Purify solanone by high vacuum distillation and store it in sealed ampules [Kohda & Sato *J Chem Soc, Chem Commun* 951 1981]. It has UV (hexane) at λ<sub>max</sub> 230nm (ε 11,800). The *semicarbazone* crystallises from aqueous EtOH or toluene with **m 160.5-161.5°**. [Johnson et al. *J Org Chem* **30** 2918 1965.]

**Sorbic acid (2,4-hexadienoic acid)** [110-44-1] **M 112.1, m 134°, pK<sup>25</sup> 4.76**. Crystallise the acid from water. Dry it air or in a desiccator over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **2** IV 1701.]

**Spermidine** [*N-(3-aminopropyl)-1,4-diaminobutane*] [124-20-9] **M 145.3, m 23-25°, b 128-131°/15mm, d<sub>4</sub><sup>20</sup> 0.918, n<sub>D</sub><sup>20</sup> 1.482, pK<sub>1</sub><sup>25</sup> 8.25, pK<sub>2</sub><sup>25</sup> 9.64, pK<sub>3</sub><sup>25</sup> 10.43**. It is a strong base with an alkylamine odour and absorbs CO<sub>2</sub> from the atmosphere. It is purified by shaking with solid K<sub>2</sub>CO<sub>3</sub> or NaOH, decanting and distilling from K<sub>2</sub>CO<sub>3</sub> in a vacuum. Store it in the dark under N<sub>2</sub>. [*Beilstein* **4** IV 1300.]

**Spermidine trihydrochloride** [334-50-9] **M 245.3, m ~250°(dec), 256-258°, for pKa see free base above**. Recrystallise the salt from dry 3% HCl in ethanol and adding dry Et<sub>2</sub>O if necessary. Filter it off rapidly and dry it in a vacuum desiccator. *Alternatively*, centrifuge the crystals off, wash them with dry Et<sub>2</sub>O and dry them in a vacuum. [*Beilstein* **4** IV 1300.]

**Spermine 4HCl** (*N,N*-bis(3-aminopropyl)-1,4-butanediamine 4HCl) [306-67-2] **M 348.2, m 313-315°**. Its pK values are similar to those of spermidine above. Purify it as for spermidine trihydrochloride above. [*Beilstein* **4** IV 1301.]

**Stearic acid (octadecanoic acid)** [57-11-4] **M 284.5, m 71.4°, 72°, b 144-145°/27mm, 383°/760mm, d<sub>4</sub><sup>20</sup> 0.911, n<sub>D</sub><sup>20</sup> 1.428, pK<sub>Est</sub> ~5.0**. Crystallise stearic acid from acetone, acetonitrile, EtOH (5 times), aqueous MeOH, ethyl methyl ketone or petroleum ether (b 60-90°), or by fractional precipitation by dissolving in hot 95% EtOH and pouring into distilled water, with stirring. The precipitate, after washing with distilled water, is dried under vacuum over P<sub>2</sub>O<sub>5</sub>. It has also been purified by zone melting and partial freezing. [Tamai et al. *J Phys Chem* **91** 541 1987, *Beilstein* **2** IV 1206.]

**Suberic acid (hexane-1,6-dicarboxylic acid)** [505-48-6] **M 174.2, m 141-142°, pK<sub>1</sub><sup>25</sup> 4.12, pK<sub>2</sub><sup>25</sup> 5.40**. Crystallise it from acetone. It sublimes at 300° without decomposition. [*Beilstein* **2** IV 2028.]

**Succinamic acid (succinic acid amide)** [638-32-4] **M 117.1, m 155°, 156-157°, pK<sup>25</sup> 4.54**. Crystallise the amide from Me<sub>2</sub>CO or H<sub>2</sub>O and dry it in a vacuum. It is not very soluble in MeOH. It is converted to succinimide above 200°. [*Beilstein* **2** H 614.]

**Succinamide** [110-14-5] **M 116.1, m 262-265°(dec)**. Crystallise it from water. [*Beilstein* **2** H 614.]

**Succinic acid** [110-15-6] **M 118.1, m 185-185.5°, pK<sub>1</sub><sup>25</sup> 4.21, pK<sub>2</sub><sup>25</sup> 5.72**. Wash it with diethyl ether. Crystallise it from acetone, distilled water, or *tert*-butanol. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub> or conc H<sub>2</sub>SO<sub>4</sub>. Also purify it by conversion to the disodium salt which, after crystallisation from boiling water (charcoal), is treated with mineral acid to regenerate the succinic acid. The acid is then recrystallised and dried in a vacuum. [*Beilstein* **2** H 606, **2** IV 1908.]

**Succinic anhydride** [108-30-5] **M 100.1, m 119-120°**. Crystallise the anhydride from redistilled acetic anhydride or CHCl<sub>3</sub>, then filter, wash with diethyl ether and dry it in a vacuum. [*Beilstein* **17** H 606, **17** V 6.]

**Succinimide** [123-56-8] **M 99.1, m 124-125°, pK<sup>25</sup> 9.62**. Crystallise the imide from EtOH (1ml/g) or water. [*Beilstein* **21** H 369, **21/9** V 438.]

**Succinonitrile** [110-61-2] **M 80.1, m 57.9°, b 108°/1mm, 267°/760mm.** Purify the nitrile by vacuum sublimation, and/or crystallisation from acetone. [Beilstein 2 H 615, 2 IV 1923.]

**D(-)-Tartaric acid** [147-71-7] **M 150.1, m 169.5-170° (2S,3S-form, natural) [ $\alpha$ ]<sub>546</sub><sup>20</sup> -15° (c 10, H<sub>2</sub>O), m 208° (2RS,3RS-form), pK<sub>1</sub><sup>25</sup> 3.03, pK<sub>2</sub><sup>25</sup> 4.46, pK<sub>3</sub><sup>25</sup> 14.4.** Crystallise the acid from distilled H<sub>2</sub>O or \*benzene/diethyl ether containing 5% of petroleum ether (b 60-80°) (1:1). Soxhlet extraction with diethyl ether has been used to remove an impurity absorbing at 265nm. It has also been crystallised from absolute EtOH/hexane and dried in a vacuum for 18 hours [Kornblum & Wade *J Org Chem* 52 5301 1987]. [Beilstein 3 IV 1229.]

**meso-Tartaric acid** [147-73-9] **M 150.1, m 139-141°, pK<sub>1</sub><sup>25</sup> 3.17, pK<sub>2</sub><sup>25</sup> 4.91.** Crystallise it from water, wash it with cold MeOH and dry it at 60° under vacuum. [Beilstein 3 IV 1218.]

**Tetra-*n*-amylammonium bromide (tetra-*n*-pentylammonium bromide)** [866-97-7] **M 378.5, m 100-101°.** Crystallise it from petroleum ether, \*benzene or acetone/ether mixtures and dry in vacuum at 40-50° for 2 days. It is used in ion-paired chromatography (Sagara et al. *J Chromatogr* 328 289 1985). [Beilstein 4 IV 677.]

**Tetra-*n*-amylammonium iodide** [2498-20-6] **M 425.5, m 135-137°.** Crystallise the iodide from EtOH and dry it at 35° under a vacuum. It has also been purified by dissolving in acetone and precipitating by adding diethyl ether, and drying at 50° for 2 days. [Beilstein 4 IV 677.]

**1,1,2,2,-Tetrabromoethane** [79-27-6] **M 345.7, f 0.0°, b 119°/15mm, 243.5°/atm, d<sub>4</sub><sup>20</sup> 2.965, n<sub>D</sub><sup>20</sup> 1.63533.** Wash it successively with conc H<sub>2</sub>SO<sub>4</sub> (three times) and H<sub>2</sub>O (three times), dry it with K<sub>2</sub>CO<sub>3</sub> and CaSO<sub>4</sub> and distil it in a vacuum or at ~760mm. [Beilstein 1 IV 162.]

**Tetra-*n*-butylammonium bromide** [1643-19-2] **M 322.4, m 119.6°.** Recrystallise the salt from \*benzene (5ml/g) at 80° by adding hot *n*-hexane (three volumes) and allowing to cool. Dry it over P<sub>2</sub>O<sub>5</sub> or Mg(ClO<sub>4</sub>)<sub>2</sub>, under vacuum. The salt is *very hygroscopic*. It can also be crystallised from ethyl acetate or dry acetone by adding diethyl ether and dried *in vacuo* at 60° for 2 days. It has been crystallised from acetone by addition of diethyl ether. It is so *hygroscopic* that all manipulations should be carried out in a dry-box. It has been purified by precipitation from a saturated solution in dry CCl<sub>4</sub> on addition of cyclohexane or by recrystallisation from ethyl acetate, then heating in vacuum to 75° in the presence of P<sub>2</sub>O<sub>5</sub>. [Symons et al. *J Chem Soc, Faraday Trans 1* 76 2251 1908.] It also recrystallises from CH<sub>2</sub>Cl<sub>2</sub>/diethyl ether and is dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. [Blau & Espenson *J Am Chem Soc* 108 1962 1986, Beilstein 4 IV 657.]

**Tetra-*n*-butylammonium chloride** [1112-67-0] **M 277.9, m 15.7°.** Crystallise the chloride from acetone by addition of diethyl ether. It is very *hygroscopic* and forms crystals with 34H<sub>2</sub>O. It is used in ion-paired chromatography (Sagara et al. *J Chromatogr* 328 289 1985). [Beilstein 4 IV 557.]

**Tetra-*n*-butylammonium hexafluorophosphate** [3109-63-5] **M 387.5, m 239-241°.** Recrystallise it from saturated EtOH/water and dry it for 10 hours in a vacuum at 70°. Also recrystallise it three times from absolute EtOH and dry it for 2 days in a drying pistol under a vacuum at boiling toluene temperature [Bedard & Dahl *J Am Chem Soc* 108 5933 1986]. It is a stable supporting electrolyte in organic solvents [Baiser in *Organic Electrochemistry* M. Dekker NY p228 1973.]

**Tetra-*n*-butylammonium hydrogen sulfate** [32503-27-8] **M 339.5, m 171-172°.** The sulfate crystallises from acetone. It has been used as a phase transfer catalyst.

**Tetra-*n*-butylammonium iodide** [311-28-4] **M 369.4, m 146°.** Crystallise the iodide from toluene/petroleum ether (see entry for the corresponding bromide), acetone, ethyl acetate, EtOH/diethyl ether, nitromethane, aqueous EtOH or water. Dry it at room temperature under a vacuum. It has also been dissolved in MeOH/acetone (1:3, 10ml/g), filtered and allowed to stand at room temperature to evaporate to *ca* half its original volume. Distilled water (1ml/g) is then added, and the precipitate is filtered off and dried. It can also

be dissolved in acetone, precipitated by adding ether and dried in a vacuum at 90° for 2 days. It has also been recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/petroleum ether or hexane, or anhydrous methanol and stored in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. [Chau & Espenson *J Am Chem Soc* **108** 1962 1986, *Beilstein* **4** IV 558.]

**Tetra-*n*-butylammonium nitrate** [1941-27-1] **M 304.5, m 119°**. Crystallise it from \*benzene (7ml/g), EtOH or EtOAc (**m** 121-122°); dry it in a vacuum over P<sub>2</sub>O<sub>5</sub> at 60° for 2 days. [*Beilstein* **4** IV 558.]

**Tetra-*n*-butylammonium perchlorate** [1923-70-2] **M 341.9°, m 210°(dec)**. Crystallise the perchlorate from EtOH, ethyl acetate, *n*-hexane or diethyl ether/acetone mixture, or hot CH<sub>2</sub>Cl<sub>2</sub>. Dry it in a vacuum at room temperature over P<sub>2</sub>O<sub>5</sub> for 24 hours. [Anson et al. *J Am Chem Soc* **106** 4460 1984, Ohst & Kochi *J Am Chem Soc* **108** 2877 1986, Collman et al. *J Am Chem Soc* **108** 2916 1986, Blau & Espenson *J Am Chem Soc* **108** 1962 1986, Gustowski et al. *J Am Chem Soc* **108** 1986, Ikezawa & Kutal *J Org Chem* **52** 3299 1987, *Beilstein* **4** IV 557.]

**Tetra-*n*-butylammonium picrate** [914-45-4] **M 490.6, m 106.5-107°**. Crystallise the picrate from EtOH. Dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **6** II 271, **4** III 292.]

**Tetra-*n*-butylammonium tetrabutylborate** (Bu<sub>4</sub>N<sup>+</sup> Bu<sub>4</sub>B<sup>-</sup>) [23231-91-6] **M 481.7, m 155°, 161.8°**. Dissolve it in MeOH or acetone, and crystallise by adding distilled water. It also crystallises from EtOH or EtOAc. Dry it in a vacuum at 70°. It has also been successively recrystallised from isopropyl ether, isopropyl ether/acetone (50:1) and isopropyl ether/EtOH (50:1) for 10 hours, then isopropyl ether/acetone for 1 hour, and dried at 65° under reduced pressure for 1 week. [Kondo et al. *J Chem Soc, Faraday Trans 1* **76** 812 1980, *Beilstein* **4** III 293, **4** III 558.]

**Tetra-*n*-butylammonium tetrafluoroborate** [429-42-5] **M 329.3, m 161.8°, 161-163°, pK<sup>25</sup> -4.9 (for HBF<sub>4</sub>)**. Recrystallise it from H<sub>2</sub>O, aqueous EtOH or from EtOAc by cooling in Dry-ice. Also recrystallise it from ethyl acetate/pentane or dry acetonitrile. Dry it at 80° under vacuum. [Detty & Jones *J Am Chem Soc* **109** 5666 1987, Hartley & Faulkner *J Am Chem Soc* **107** 3436 1985.] The *acetate* has **m** 118±2° (from BuCl), the *bromide* has **m** 118° (from EtOAc), and the *nitrate* has **m** 120° (from \*C<sub>6</sub>H<sub>6</sub>). [Witschonka & Kraus *J Am Chem Soc* **69** 2472 1947, Wheeler & Sandstedt *J Am Chem Soc* **77** 2024 1955, *Beilstein* **4** IV 558.]

**1,1,2,2-Tetrachloro-1,2-difluoroethane** [72-12-0] **M 203.8, f 26.0°, b 92.8°/760mm, d<sub>4</sub><sup>25</sup> 1.6252, n<sub>D</sub><sup>25</sup> 1.4130**. Purify it as for trichlorotrifluoroethane. [*Beilstein* **1** III 165, **1** III 146.]

***sym*-Tetrachloroethane** [79-34-5] **M 167.9, b 62°/100mm, 146.2°/atm, d<sub>4</sub><sup>20</sup> 1.588, n<sub>D</sub><sup>15</sup> 1.49678**. Stir the ethane, on a steam-bath, with conc H<sub>2</sub>SO<sub>4</sub> until a fresh portion of acid remains colourless. The organic phase is then separated, distilled in steam, dried (CaCl<sub>2</sub> or K<sub>2</sub>CO<sub>3</sub>), and fractionally distilled in a vacuum. [*Beilstein* **1** IV 144.]

**Tetrachloroethylene** [127-18-4] **M 165.8, b 62°/80mm, 121.2°, d<sub>4</sub><sup>15</sup> 1.63109, d<sub>4</sub><sup>20</sup> 1.623, n<sub>D</sub><sup>15</sup> 1.50759, n<sub>D</sub><sup>20</sup> 1.50566**. It decomposes under similar conditions to CHCl<sub>3</sub>, to give phosgene and trichloroacetic acid. Inhibitors of this reaction include EtOH, diethyl ether and thymol (effective at 2-5 ppm). Tetrachloroethylene should be distilled under a vacuum (to avoid formation of phosgene) and stored in the dark out of contact with air. It can be purified by washing with 2M HCl until the aqueous phase no longer becomes coloured, then with water, drying with Na<sub>2</sub>CO<sub>3</sub>, Na<sub>2</sub>SO<sub>4</sub>, CaCl<sub>2</sub> or P<sub>2</sub>O<sub>5</sub>, and fractionally distilling just before use. 1,1,2-Trichloroethane and 1,1,1,2-tetrachloroethane can be removed by counter-current extraction with EtOH/water. [*Beilstein* **1** IV 715.]

**Tetracosane (C<sub>24</sub>)** [646-31-1] **M 338.7, m 54°, b 243-244°/15mm**. Crystallise it from diethyl ether and/or distil it under high vacuum. [*Beilstein* **1** IV 578.]

**Tetracosanoic (lignoceric) acid** [557-59-5] **M 368.7, m 84°, 87.5-88°, pK<sub>Est</sub> ~5.0**. Crystallise the acid from acetic acid, Me<sub>2</sub>CO, toluene, petroleum ether/Me<sub>2</sub>CO or \*C<sub>6</sub>H<sub>6</sub>/Me<sub>2</sub>CO. [*Beilstein* **2** IV 1301.]

**Tetracyanoethylene** [670-54-2] **M 128.1, m 199-200°(sealed tube).** Rerystallise it from chlorobenzene, dichloroethane, or dichloromethane [Hall et al. *J Org Chem* **52** 5528 1987]. Store it at 0° in a desiccator over NaOH pellets. (It slowly evolves HCN on exposure to moist air **CARE.**) It can also be sublimed at 120° under vacuum. Also purify it by repeated sublimation at 120-130°/0.5mm. [Frey et al. *J Am Chem Soc* **107** 748 1985, Traylor & Miksztal *J Am Chem Soc* **109** 2778 1987, Fatiadi *Synthesis* 249 1986, *Synthesis* 749 1967, *Beilstein* **2** IV 1245.]

**Tetradecane (C14)** [629-59-4] **M 198.4, m 6°, b 122°/10mm, 252-254°, d<sub>4</sub><sup>20</sup> 0.763, n<sub>D</sub><sup>20</sup> 1.429.** Wash it successively with 4M H<sub>2</sub>SO<sub>4</sub> and water. Dry it over MgSO<sub>4</sub> and distil it several times under reduced pressure [Poë et al. *J Am Chem Soc* **108** 5459 1986]. It is used as a standard in gas chromatography. [*Beilstein* **1** H 171, **1** IV 520.]

**1-Tetradecanol** [112-72-1] **M 214.4, m 39-39.5°, b 160°/10mm, 170-173°/20mm.** Crystallise the alcohol from aqueous EtOH. It has also been purified by zone melting. [*Beilstein* **1** IV 1864.]

**Tetradecyl ether (di-tetradecyl ether)** [5412-98-6] **M 410.7, m 43.5°, d<sub>4</sub><sup>43</sup> 0.8117.** Distil the ether under a vacuum and then crystallise it repeatedly from MeOH/\*benzene. It also crystallises from MeOH alone (m 31.2°, 33°, and 44.4° also reported), or Me<sub>2</sub>CO (m 43.5°). [Di Giacomo & Smyth *J Am Chem Soc* **78** 2027 1956, *Beilstein* **1** IV 1865.]

**Tetradecyltrimethylammonium bromide (myristyl trimethylammonium bromide)** [1119-97-7] **M 336.4, m 244-245°, 244-249°.** Crystallise the bromide from Me<sub>2</sub>CO or a mixture of Me<sub>2</sub>CO and >5% MeOH or Me<sub>2</sub>CO and EtOH. Wash it with diethyl ether and dry it in a vacuum oven at 60°. It is a cationic detergent. Its solubility is 1g/5g H<sub>2</sub>O. [Dearden & Wooley *J Phys Chem* **91** 2404 1987, Shelton et al. *J Am Chem Soc* **68** 754 1946, *Beilstein* **4** III 419, **4** IV 813.]

**Tetraethoxymethane** See tetraethyl orthocarbonate below.

**Tetraethylammonium bromide** [71-91-0] **M 210.2, m 269°(dec), 284°(dec).** Recrystallise the bromide from EtOH, CHCl<sub>3</sub> or diethyl ether, or recrystallise it from acetonitrile and dry it over P<sub>2</sub>O<sub>5</sub> under reduced pressure for several days. It also recrystallises from EtOH/diethyl ether (1:2), EtOAc, water or boiling MeOH/acetone (1:3) or by adding an equal volume of acetone and allowing to cool. Dry it at 100° *in vacuo* for 12 days, and store over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **4** IV 332.]

**Tetraethylammonium chloride hydrate** [68696-18-4 (H<sub>2</sub>O), 56-34-8 (anhydrous)] **M 165.7, m dec >200°.** Crystallise the chloride from EtOH by adding diethyl ether, from warm water by adding EtOH and diethyl ether, from dimethylacetamide or from CH<sub>2</sub>Cl<sub>2</sub> by addition of diethyl ether. Dry it over P<sub>2</sub>O<sub>5</sub> in vacuum for several days. It also crystallises from acetone/CH<sub>2</sub>Cl<sub>2</sub>/hexane (2:2:1) [Blau & Espenson *J Am Chem Soc* **108** 1962 1986, White & Murray *J Am Chem Soc* **109** 2576 1987]. [*Beilstein* **4** IV 332.]

**Tetraethylammonium hexafluorophosphate** [429-07-2] **M 275.2, m >300°, 331°(dec), pK<sub>1</sub><sup>25</sup> ~0.5, pK<sub>2</sub><sup>25</sup> 5.12 (for fluorophosphoric acid H<sub>2</sub>PO<sub>3</sub>F).** Dissolve the salt (0.8g) in hot H<sub>2</sub>O (3.3ml) and cool to crystallise. Yield of prisms is 0.5g. Its solubility in H<sub>2</sub>O is 8.1g/L at 19° [Lange & Müller *Chem Ber* **63** 1067 1930]. [*Beilstein* **4** III 199.]

**Tetraethylammonium iodide** [68-05-3] **M 257.2, m 302°, >300°(dec).** Crystallise the iodide from acetone/MeOH, EtOH/water, dimethylacetamide or ethyl acetate/EtOH (19:1). Dry it under a vacuum at 50° and store it over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **4** IV 332.]

**Tetraethylammonium perchlorate** [2567-83-1] **M 229.7, m 345°(dec).** Crystallise the perchlorate repeatedly from water, aqueous MeOH, acetonitrile or acetone, and dry it at 70° under a vacuum for 24 hours. [Cox et al. *J Am Chem Soc* **106** 5965 1984, Liu et al. *J Am Chem Soc* **108** 1740 1986, White & Murray *J Am Chem Soc* **109** 2576 1987.] It has also been crystallised twice from ethyl acetate/95% EtOH (2:1) [Lexa et al. *J Am Chem Soc*

109 6464 1987]. [*Beilstein* 4 IV 332.]

**Tetraethylammonium picrate** [741-03-7] **M 342.1, m >300°(dec)**. Purify it by successive crystallisations from water or 95% EtOH followed by drying in vacuum at 70°. [*Beilstein* 4 IV 332.]

**Tetraethylammonium tetrafluoroborate** [429-06-1] **M 217.1, m 235°, 356-367°, 275-277°, 289-291°, pK<sup>25</sup> -4.9 (for HBF<sub>4</sub>)**. Dissolve the salt in hot MeOH, filter and add Et<sub>2</sub>O. It is soluble in ethylene chloride [Thompson & Kraus *J Am Chem Soc* 69 1016 1947, Wheeler & Sandstadt 77 2025 1955]. It has also been recrystallised three times from a mixture of ethyl acetate/hexane (5:1) or MeOH/petroleum ether, then stored at 95° for 48 hours under vacuum [Henry & Faulkner *J Am Chem Soc* 107 3436 1985, Huang et al. *Anal Chem* 58 2889 1986]. It is used as a supporting electrolyte. [*Beilstein* 4 IV 333.]

**Tetraethylammonium tetraphenylborate** [12099-10-4] **M 449.4**. Recrystallise the borate from aqueous acetone. Dry it in a vacuum oven at 60° for several days. *Similarly for the propyl and butyl homologues*. [*Beilstein* 4 IV 333.]

**Tetraethylene glycol dimethyl ether** [143-24-8] **M 222.3, b 105°/1mm, d<sub>4</sub><sup>20</sup> 1.010, n<sub>D</sub><sup>20</sup> 1.435**. Stand the ether over CaH<sub>2</sub>, LiAlH<sub>4</sub> or sodium, and distil it when required. [*Beilstein* 1 IV 2404.]

**Tetraethylenepentamine** [112-57-2] **M 189.3, b 169-171°/0.05mm, d<sub>4</sub><sup>20</sup> 0.999, n<sub>D</sub><sup>20</sup> 1.506, pK<sub>1</sub><sup>25</sup> 2.98, pK<sub>2</sub><sup>25</sup> 4.72, pK<sub>3</sub><sup>25</sup> 8.08, pK<sub>4</sub><sup>25</sup> 9.10, pK<sub>5</sub><sup>25</sup> 9.68**. Distil the amine under vacuum. Also purify *via* its penta hydrochloride, nitrate or sulfate. Jonassen, Frey and Schaafsma [*J Phys Chem* 61 504 1957] cooled a solution of 150g of the base in 300ml of 95% EtOH, and added dropwise 180ml of conc HCl, keeping the temperature below 20°. The white precipitate was filtered off, crystallised three times from EtOH/water, then washed with diethyl ether and dried by suction. Reilley & Holloway [*J Am Chem Soc* 80 2917 1958], starting with a similar solution cooled to 0°, added slowly (keeping the temperature below 10°) a solution of 4.5g-moles of HNO<sub>3</sub> in 600ml of aqueous 50% EtOH (also cooled to 0°). The precipitate was filtered by suction, recrystallised five times from aqueous 5% HNO<sub>3</sub>, then washed with acetone and absolute EtOH and dried at 50°. [For purification *via* the sulfate see Reilley and Vavoulis (*Anal Chem* 31 243 1959), and for an additional purification step using the Schiff base with benzaldehyde see Jonassen et al. *J Am Chem Soc* 79 4279 1957]. [*Beilstein* 4 IV 1244.]

**Tetraethyl 1,1,2,2-ethanetetracarboxylate** [632-56-4] **M 318.3, m 73-74°**. Recrystallise the ester twice from EtOH by cooling to 0°. [Mochizuki et al. *Bull Chem Soc Jpn* 64 1750 1991, Weinges et al. *Angew Chem* 93 1008 1981, *Beilstein* 2 IV 2415.]

**Tetraethyl orthocarbonate (ethyl orthocarbonate, tetraethoxy ethane)** [78-09-1] **M 192.3, b 59.6-60°/14mm, 158°/atm, 159°/atm, 160-161°/atm, d<sub>4</sub><sup>20</sup> 0.9186, n<sub>D</sub><sup>20</sup> 1.3932**. Likely impurities are hydrolysis products. Shake the orthocarbonate with brine (saturated NaCl, dilute with a little Et<sub>2</sub>O if amount of material is small) and dry (MgSO<sub>4</sub>). The organic layer is filtered off and evaporated, and the residue is distilled through a helices packed fractionating column with a total reflux partial take-off head. All distillations can be done at atmospheric pressure in an inert atmosphere (e.g. N<sub>2</sub>). [Roberts & McMahon *Org Synth Coll Vol* IV 457 1963, Connolly & Dyson *J Chem Soc* 828 1937, Tieckelmann & Post *J Org Chem* 13 266 1948, for review see Kantlehner et al. *Justus Liebig's Ann Chem* 507 207 1982, *Beilstein* 3 IV 6.]

**2,2,3,3-Tetrafluoropropan-1-ol** [76-37-9] **M 132.1, b 106-106.5°/~760mm, pK<sup>25</sup> 12.74**. Tetrafluoro-*n*-propanol (450ml) is added to a solution of NaHSO<sub>3</sub> (2.25g) in water (90ml), shaken vigorously and set aside for 24 hours. The fraction distilling at or above 99° is refluxed for 4 hours with 5-6g of KOH and rapidly distilled, followed by a final fractional distillation. [Kosower & Wu *J Am Chem Soc* 83 3142 1961.] *Alternatively*, shake the alcohol with alumina for 24 hours, dry it overnight with anhydrous K<sub>2</sub>CO<sub>3</sub> and distil it, taking the middle fraction (**b** 107-108°). [*Beilstein* 1 IV 2438.]

**Tetra-*n*-heptylammonium bromide** [4368-51-8] **M 490.7, m 88.9-89.1°**. Crystallise the bromide from *n*-hexane, then dry it in a vacuum oven at 70°. [Goodrich et al. *J Am Chem Soc* 72 4412 1950, *Beilstein* 4 IV 736.]

**Tetra-*n*-heptylammonium iodide** [3535-83-9] **M 537.7, m 102-103°**. Recrystallise the iodide from EtOH or aqueous EtOH. [Eriksen et al. *J Org Chem* **25** 849 1960, *Beilstein* **4** IV 736 for triheptylamine.]

**Tetra-*n*-hexylammonium bromide** [4328-13-6] **M 434.6, m 99-100°**. Wash the bromide with ether, and dry it in a vacuum at room temperature for 3 days.

**Tetra-*n*-hexylammonium chloride** [5922-92-9] **M 390.1**. Crystallise the chloride from EtOH.

**Tetra-*n*-hexylammonium iodide** [2138-24-1] **M 481.6, m 99-101°, 102-103°**. Wash the iodide with diethyl ether and dry it at room temperature *in vacuo* for 3 days. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>. [Eriksen et al. *J Org Chem* **25** 849 1960, *Beilstein* **4** IV 711 for trihexylamine.]

**Tetrahexylammonium perchlorate** [4656-81-9] **M 454.1, m 104-106°**. Crystallise the salt from acetone and dry it *in vacuo* at 80° for 24 hours.

**Tetrakis(dimethylamino)ethylene** [996-70-3] **M 300.2, b 60°/1mm, d<sub>4</sub><sup>20</sup> 0.861, n<sub>D</sub><sup>20</sup> 1.4817, pK<sub>Est(1)</sub><0, pK<sub>Est(2)</sub><0, pK<sub>Est(3)</sub>~1.5, pK<sub>Est(4)</sub>~5.1**. Impurities include tetramethylurea, dimethylamine, tetramethylethane-diamine and tetramethyloxamide. It is washed with water while being flushed with nitrogen to remove dimethylamine, dried over molecular sieves, then passed through a silica gel column (previously activated at 400°) under nitrogen. De-gas it in a vacuum line by distillation from a trap at 50° to one at -70°. Finally, it is stirred over sodium-potassium alloy for several days. [Holroyd et al. *J Phys Chem* **89** 4244 1985, Wiberg *Angew Chem Int Ed Engl* **7** 766 1968, *Beilstein* **4** IV 167.]

**Tetramethylammonium bromide** [64-20-0] **M 154.1, sublimes with dec >230°**. Crystallise the bromide from EtOH, EtOH/diethyl ether, MeOH/acetone, water or from acetone/MeOH (4:1) by adding an equal volume of acetone. It is dried at 110° under reduced pressure or at 140° for 24 hours. [*Beilstein* **4** IV 145.]

**Tetramethylammonium chloride** [75-57-0] **M 109.6, m >230°(dec)**. Crystallise the chloride from EtOH, EtOH/CHCl<sub>3</sub>, EtOH/diethyl ether, acetone/EtOH (1:1), isopropanol or water. Traces of the free amine can be removed by washing with CHCl<sub>3</sub>. [*Beilstein* **4** IV 145.]

**Tetramethylammonium hydroxide (5H<sub>2</sub>O)** [10424-65-4 (5H<sub>2</sub>O), 75-59-2 (aqueous solution)] **M 181.2, m 63°, 65-68°**. It is freed from chloride ions by passage through an ion-exchange column (e.g. Amberlite IRA-400, prepared in its OH<sup>-</sup> form by passing 2M NaOH until the effluent is free from chloride ions, then washed with distilled H<sub>2</sub>O until neutral). A modification, to obtain carbonate-free hydroxide, uses the method of Davies and Nancollas [*Nature* **165** 237 1950]. [*Beilstein* **4** IV 145.]

**Tetramethylammonium hexafluorophosphate** [558-32-7] **M 219.1, m >300°, d<sub>4</sub><sup>25</sup> 1.617, pK<sub>1</sub><sup>25</sup> ~0.5, pK<sub>2</sub><sup>25</sup> 5.12 (for fluorophosphoric acid H<sub>2</sub>PO<sub>3</sub>F)**. The salt (0.63g) is recrystallised from boiling H<sub>2</sub>O (76ml), yielding pure (0.45) Me<sub>4</sub>N.PF<sub>6</sub> after drying at 100°. It is a good supporting electrolyte. [Lange & Müller *Chem Ber* **63** 1067 1930, *Beilstein* **4** III 110.]

**Tetramethylammonium iodide** [75-58-1] **M 201.1, m >230°(dec)**. Crystallise the iodide from water or 50% EtOH, EtOH/diethyl ether, ethyl acetate, or from acetone/MeOH (4:1) by adding an equal volume of acetone. Dry it in a vacuum desiccator. [*Beilstein* **4** IV 145.]

**Tetramethylammonium nitrate** [1941-24-8] **M 136.2, m >300°, 410°**. Recrystallise the nitrate from EtOH and dry at 110° in an air oven. [Coats & Taylor *J Chem Soc* 1498 1936, *Beilstein* **4** III 113, **4** IV 147.]

**Tetramethylammonium perchlorate** [2537-36-2] **M 123.6, m >300°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Crystallise it twice from H<sub>2</sub>O and dry it at 110° in an air oven. It is insoluble in most organic solvents. [Mead et al. *J Chem Soc* 1210 1933, Coats & Taylor *J Chem Soc* 1498 1936.]

**Tetramethylammonium tetraphenylborate** [15525-13-0] **M 393.3**. Recrystallise it from acetone, acetone/CCl<sub>4</sub> and from acetone/1,2-dichloroethane. Dry it over P<sub>2</sub>O<sub>5</sub> in a vacuum, or in a vacuum oven at 60° for several days. [*Beilstein* 4 IV 145.]

***N,N,N',N'*-Tetramethylethylenediamine (TMEDA, TMED)** [110-18-9] **M 116.2, m -55°, b 122°, d<sub>4</sub><sup>20</sup> 1.175, n<sub>D</sub><sup>25</sup> 1.4153, pK<sub>1</sub><sup>25</sup> 5.90, pK<sub>2</sub><sup>25</sup> 9.14**. Dry TMEDA partially with molecular sieves (Linde type 4A), then distil it in a vacuum from butyl lithium. This treatment removes all traces of primary and secondary amines and water. [Hay et al. *J Chem Soc, Faraday Trans 1* 68 1 1972.] Or dry it with KOH pellets, reflux for 2 hours with one-sixth its weight of *n*-butyric anhydride (to remove primary and secondary amines) and fractionally distil it. Reflux it with fresh KOH, and distil it under nitrogen. [Cram & Wilson *J Am Chem Soc* 85 1245 1963.] It was also distilled from Na. Store it sealed under N<sub>2</sub>. The *dipicrate* has **m 263°(dec)**. [*Beilstein* 4 H 250, 4 I 415, 4 II 690, 4 III 512, 4 IV 1172.]

**Tetramethylethylenediamine dihydrochloride** [7677-21-8] **M 198.2, m ~300°**. Crystallise the salt from 98% EtOH/conc HCl. It is *hygroscopic*. [Knorr *Chem Ber* 37 3510 1904, *Beilstein* 4 IV 1172.]

**1,1,3,3-Tetramethylguanidine** [80-70-6] **M 115.2, b 159-160°, d<sub>4</sub><sup>20</sup> 0.917 n<sub>D</sub><sup>20</sup> 1.470, pK<sub>25</sub> 13.6**. Reflux it over granulated BaO, then fractionally distil it. Protect it from CO<sub>2</sub>. [*Beilstein* 4 IV 227.]

**Tetramethyl orthocarbonate (methyl orthocarbonate, tetramethoxy methane)** [1850-14-2] **M 136.2, m -5.6°, -5°, -2°, b 113.5°/760mm, 113.5-114°/755mm, 112-114°/atm, d<sub>4</sub><sup>20</sup> 1.0202, n<sub>D</sub><sup>20</sup> 1.3860**. Purify it in the same way as for tetraethyl orthocarbonate. [Smith *Acta Chem Scand* 10 1006 1956, Tiekemann & Post *J Org Chem* 13 266 1948, Kantlehner et al. *Synthesis* 73 1977, *Beilstein* 3 IV 4.]

**2,6,10,14-Tetramethylpentadecane (pristane, norphytane)** [1921-70-6] **M 268.5, b 68° (bath temp)/0.004mm, 158°/10mm, 296°/atm, d<sub>4</sub><sup>20</sup> 0.7827, n<sub>D</sub><sup>20</sup> 1.4385**. Purify pristane by shaking it with conc H<sub>2</sub>SO<sub>4</sub> (**care**, if amount of pristane is too small then it should be diluted with petroleum ether *not* Et<sub>2</sub>O which is quite soluble in H<sub>2</sub>SO<sub>4</sub>), then H<sub>2</sub>O (**care**, as it may heat up in contact with conc H<sub>2</sub>SO<sub>4</sub>), dry (MgSO<sub>4</sub>), evaporate and distil it over Na. [Sørensen & Sørensen *Acta Chem Scand* 3 939 1949, *Beilstein* 1 III 570.]

**Tetramethylthiuram disulfide [bis-(dimethylthiocarbamyl)disulfide, Thiram]** [137-26-8] **M 240.4, m 146-148°, 155-156°**. Crystallise thiram (three times) from boiling CHCl<sub>3</sub>, then recrystallise it from boiling CHCl<sub>3</sub> by adding EtOH dropwise to initiate crystallisation, and allow it to cool. Finally, it is precipitated from cold CHCl<sub>3</sub> by adding EtOH (which retains the monosulfide in solution). [Ferington & Tobolsky *J Am Chem Soc* 77 4510 1955, *Beilstein* 4 IV 242.]

**1,1,3,3-Tetramethyl urea** [632-22-4] **M 116.2, f -1.2°, b 175.2°/760mm, d<sub>4</sub><sup>20</sup> 0.969, n<sub>D</sub><sup>20</sup> 1.453**. Dry it over BaO and distil it under nitrogen. It denatures proteins in H<sub>2</sub>O. [Elbaum & Herskovits *Biochemistry* 13 1268 1974, Kane *Anal Biochem* 53 350 1973, *Beilstein* 4 IV 225.]

**Tetranitromethane** [509-14-8] **M 196.0, m 14.2°, b 46°/36mm, 21-23°/23mm, 126°/760mm, d<sub>4</sub><sup>20</sup> 1.640, n<sub>D</sub><sup>20</sup> 1.438**. Shake tetranitromethane with dilute NaOH, wash (H<sub>2</sub>O), steam distil, dry with Na<sub>2</sub>SO<sub>4</sub> and fractionally crystallise it by partial freezing. The melted crystals are dried with MgSO<sub>4</sub> and fractionally distilled under reduced pressure. *Alternatively*, shake it with a large volume of dilute NaOH until no absorption attributable to the *aci*-nitro anion (from mono- di- and tri- nitromethanes) is observable in the water. Then wash it with distilled water, and distil it at room temperature by passing a stream of air or nitrogen through the liquid and condensing it in a trap at -80°. It can be dried with MgSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, fractionally crystallised from the melt, and fractionally distilled under reduced pressure. [Liang *Org Synth Coll Vol III* 803 1955, *Beilstein* 4 H 80, 4 I 21, 4 II 45, 4 III 116, 4 IV 107.] **Potentially explosive (when impure e.g. with toluene), toxic, carcinogenic.**

**Tetra-*n*-propylammonium bromide** [1941-30-6] **M 266.3, m >280°(dec)**. Crystallise it from ethyl acetate/EtOH (9:1), acetone or MeOH. Dry it at 110° under reduced pressure. [*Beilstein* 4 IV 471.]

**Tetra-*n*-propylammonium iodide** [631-40-3] M 313.3, m >280°(dec). Purify the iodide by crystallising it from EtOH, EtOH/diethyl ether (1:1), EtOH/water or aqueous acetone. Dry it at 50° under a vacuum and store it over P<sub>2</sub>O<sub>5</sub> in a vacuum desiccator. Keep it away from light. [Beilstein 4 IV 472.]

**Tetra-*n*-propylammonium perchlorate** [15780-02-6] M 285.8, m 238-240°, 239-241°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>). Purify it by recrystallisation from H<sub>2</sub>O or MeCN/H<sub>2</sub>O (1:4.v/v), and dry it in an oven at 60° for several days, or in a vacuum over P<sub>2</sub>O<sub>5</sub> at 100°. [Walden & Hilgert *Z Phys Chem* 165 245 1933, Walden & Birr *Z Phys Chem* 144 281 1929, Walden & Busch *Z Phys Chem* 140 97 1929, Beilstein 4 II 628.]

**Thioacetamide** [62-55-5] M 75.1, m 112-113°, pK<sup>25</sup> 13.4. Crystallise the amide from absolute diethyl ether or \*benzene. Dry it at 70° in a vacuum and store it over P<sub>2</sub>O<sub>5</sub> at 0° under nitrogen. (It develops an obnoxious odour on storage, and the absorption at 269nm decreases, hence it should be freshly recrystallised before use). [Beilstein 2 IV 565.]

**Thiodiglycollic acid (2,2'-dithioacetic acid)** [123-93-3] M 150.2, m 129°, pK<sub>1</sub><sup>25</sup> 3.15 (3.24), pK<sub>2</sub><sup>25</sup> 4.13 (4.56). Crystallise the acid from water. [Beilstein 3 IV 612.]

**3,3'-Thiodipropionic acid (bis[2-carboxyethyl]sulfide)** [111-17-1] M 178.2, m 134°, pK<sub>1</sub><sup>25</sup> 3.84, pK<sub>2</sub><sup>25</sup> 4.66. Crystallise the sulfide from water (very soluble at 100°, but 3.7% at 26°). Antioxidant. [Beilstein 3 IV 735.]

**Thioformamide** [115-08-2] M 61.0, m 29°, 32.0-33.8°, pK<sub>Est</sub> ~12.4. Crystallise thioformamide from EtOAc, Et<sub>2</sub>O or ether/petroleum ether. The *monohydrate* is a yellow oil soluble in many organic solvents. Its UV has λ<sub>max</sub> at 263nm (ε 2500) in MeOH. [Erlenmyer & Menzi *Helv Chim Acta* 31 2071 1948.] Alternatively, dissolve it in Et<sub>2</sub>O to separate it from any formamide and/or polymers, filter, evaporate and recrystallise the residue from EtOAc at Dry-ice temperature [Londergan et al. *J Am Chem Soc* 75 4456 1953]. Store it in Et<sub>2</sub>O solution over P<sub>2</sub>O<sub>5</sub>. [Cousineau & Secrist *J Org Chem* 44 4351 1979, Beilstein 2 H 95, 2 I 39, 2 III 128, 2 IV 92.]

**Thioglycollic acid** [68-11-1] M 92.1, b 95-96°/8mm, d<sub>4</sub><sup>20</sup> 1.326, n<sub>D</sub><sup>20</sup> 1.505, pK<sub>1</sub><sup>25</sup> 3.42, pK<sub>2</sub><sup>25</sup> 10.20. Mix the acid with an equal volume of \*benzene; the \*benzene is then distilled off to dehydrate the acid. After heating to 100° to remove most of the \*benzene, the residue is distilled under vacuum and stored in sealed ampoules at 3°. [Eshelman et al. *Anal Chem* 22 844 1960, Beilstein 3 IV 1130.]

(±)-**Thiomalic (mercaptosuccinic) acid** [70-49-5] M 150.2, m 153-154°, pK<sub>1</sub><sup>25</sup> 3.64 (3.17), pK<sub>2</sub><sup>25</sup> 4.64 (4.67), pK<sub>3</sub><sup>25</sup> 10.37 (10.52). Dissolve the acid in water and extract it several times with diethyl ether to remove impurities. The aqueous solution gave the acid on freeze-drying. [Beilstein 3 IV 472.]

**Thiosemicarbazide** [79-19-6] M 91.1, m 181-183°, pK<sub>1</sub><sup>25</sup> 1.88, pK<sub>2</sub><sup>25</sup> 12.81. Crystallise thiosemicarbazide from H<sub>2</sub>O (solubility is 20.3% w/w at 80°). The *hydrochloride* has m 190-191°(dec, 184° also reported). It forms salts with heavy metals. [Beilstein 3 H 195, 3 I 79, 3 II 134, 3 III 315, 3 IV 374.]

**Thiourea** [62-56-6] M 76.1, m 179°, pK<sup>20</sup> -1.19 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise thiourea from absolute EtOH, MeOH, acetonitrile or water. Dry it under vacuum over H<sub>2</sub>SO<sub>4</sub> at room temperature. [Beilstein 3 IV 342.]

**Tiglic acid (trans-2,3-dimethylacrylic acid)** [80-59-1] M 100.1, m 63.5-64°, b 198.5°, 95°/11mm, pK<sup>18</sup> 4.96. Crystallise it from water. It is steam volatile and is soluble in organic solvents. [Beilstein 2 IV 1552.]

**trans-Traumatic acid (2-dodecene-1,12-dioic acid)** [6402-36-4] M 228.3, m 165-166°, 150-160°/0.001mm, pK<sub>Est(1)</sub> ~4.2, pK<sub>Est(2)</sub> ~4.6. Crystallise the acid from EtOH, acetone or glyme. The *bis-4-phenylphenacyl ester* has m 144-145° (from EtOH). [Beilstein 2 III 1978, 2 IV 2279.]

**1,2,3-Triaminopropane trihydrochloride** [free base 21291-99-6] M 198.7, m 250° (sintering at 100°), pK<sub>1</sub><sup>20</sup> 3.72, pK<sub>2</sub><sup>20</sup> 7.95, pK<sub>3</sub><sup>20</sup> 9.59. Crystallise the trihydrochloride from EtOH or H<sub>2</sub>O. The *free base*

decomposes at 190°/760mm, but has **b 92-93°/9mm** without decomposition. [*Beilstein* 4 H 274, 4 III 630.]

**Tribromochloromethane** [594-15-0] **M 287.2, m 55°, b 158-159.5°/~760mm, 160°/~760mm.** Melt it, wash it with aqueous Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, dry it with BaO and fractionally crystallise from its melt. It also crystallises from EtOH and distils at atmospheric pressure. [*Beilstein* 1 H 68, 1 II 35, 1 III 91, 1 IV 85.]

**Tri-*n*-butylamine (TBA)** [102-82-9] **M 185.4, b 68°/3mm, 120°/44mm, d<sub>4</sub><sup>20</sup> 0.7788, n<sub>D</sub><sup>20</sup> 1.4294, pK<sup>25</sup> 9.93.** Purify the amine by fractional distillation from sodium under reduced pressure. Pegolotti and Young [*J Am Chem Soc* 83 3251 1961] heated the amine overnight with an equal volume of acetic anhydride, in a steam bath. The amine layer was separated and heated with water for 2 hours on the steam bath (to hydrolyse any remaining acetic anhydride). The solution was cooled, solid K<sub>2</sub>CO<sub>3</sub> was added to neutralise any acetic acid that had been formed, and the amine was separated, dried (K<sub>2</sub>CO<sub>3</sub>) and distilled at 44mm pressure. Davis and Nakshbendi [*J Am Chem Soc* 84 2085 1926] treated the amine with one-eighth of its weight of benzenesulfonyl chloride in aqueous 15% NaOH at 0-5°. The mixture was shaken intermittently and allowed to warm to room temperature. After a day, the amine layer was washed with aqueous NaOH, then water and dried with KOH. (This treatment removes primary and secondary amines.) It was further dried with CaH<sub>2</sub> and distilled under vacuum. [*Beilstein* 4 IV 554.]

**Tri-*n*-butylammonium hydrobromide** [37026-85-0] **M 308.3, m 75.2-75.9°.** Crystallise the hydrobromide from ethyl acetate. [*Beilstein* 4 H 157, 4 III 292, 4 IV 555.]

**Tri-*n*-butylammonium nitrate** [33850-87-2] **M 304.5.** Crystallise the nitrate from mixtures of *n*-hexane and acetone (95:5). Dry it over P<sub>2</sub>O<sub>5</sub> in a vacuum. [*Beilstein* 2 IV 554.]

**Tri-*n*-butylammonium perchlorate** [14999-66-7] **M 285.5.** Recrystallise the perchlorate from *n*-hexane. (Potentially explosive.) [*Beilstein* 2 IV 554.]

**Tricarballic acid (propane-1,2,3-tricarboxylic acid)** [99-14-9] **M 176.1, m 166°, pK<sub>1</sub><sup>25</sup> 3.47, pK<sub>2</sub><sup>25</sup> 4.54, pK<sub>3</sub><sup>25</sup> 5.89.** Crystallise the acid from diethyl ether. [*Beilstein* 2 IV 2366.]

**Trichloroacetamide** [594-65-0] **M 162.4, m 139-141°, b 238-240°.** Its solution in xylene is dried with P<sub>2</sub>O<sub>5</sub>, then fractionally distilled. [*Beilstein* 2 IV 520.]

**Trichloroacetic acid** [76-03-9] **M 163.4, m 59.4-59.8°, pK<sup>25</sup> 0.51.** Purify this strong acid (care) by fractional crystallisation from its melt, then crystallise it repeatedly from dry \*benzene and store it over conc H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. It can also be crystallised from CHCl<sub>3</sub> or cyclohexane, and dried over P<sub>2</sub>O<sub>5</sub> or Mg(ClO<sub>4</sub>)<sub>2</sub> in a vacuum desiccator. Trichloroacetic acid can be fractionally distilled under reduced pressure from MgSO<sub>4</sub>. Layne, Jaffé and Zimmer [*J Am Chem Soc* 85 435 1963] dried trichloroacetic acid in \*benzene by distilling off the \*benzene-water azeotrope, then crystallised the acid from the remaining \*benzene solution. Manipulations should be carried out under N<sub>2</sub>. [*Toxic vapours, use a well ventilated fume cupboard.*] [*Beilstein* 2 IV 508.]

**Trichloroacetonitrile** [545-06-2] **M 144.4, m -42°, -44°, b 84.6°/741mm, 85.8-86°/764mm, 85.7°/760mm, d<sub>4</sub><sup>20</sup> 1.441, n<sub>D</sub><sup>25</sup> 1.4409.** It is prepared by mixing trichloroacetamide (150g, obtained from ethyl trichloroacetate and ammonia) with an equal weight of P<sub>2</sub>O<sub>5</sub> and heating at ~200° in an oil bath, and a further amount of P<sub>2</sub>O<sub>5</sub> is added before distilling the nitrile off. The distillate is then carefully fractionated, and the fraction **b 85.8-86°/764mm** is collected (yield 70-75%), and should contain <0.02% of amide: *cf* IR (film) has a nitrile band at 2250 cm<sup>-1</sup>, and is almost free from bands at ~1735 (CO) and >2502 (NH) cm<sup>-1</sup>. The liquid shows no signs of change after several months, even with traces of moisture, provided that it is kept in away from light. [Davies & Jenkins *J Chem Soc* 2374 1954, and references therein, Carpenter *J Org Chem* 27 2085 1962, *Beilstein* 2 H 212, 2 I 95, 2 II 201, 2 III 477, 2 IV 524]. **It is lachrymatory, a skin and eye irritant, and has been used as a fumigant.**

It is a useful reagent for selectively esterifying phosphoric acid, e.g. with benzyl alcohol in the presence of Et<sub>2</sub>N it gives monobenzylphosphate with H<sub>2</sub>PO<sub>4</sub> [Cramer & Baldauf *Chem Ber* 92, 370 1959], and converts a monophosphoric ester into a symmetrical pyrophosphate in the presence of pyridine [Cramer et al. *Justus*

*Liebigs Ann Chem* **654** 180 1962]. It is readily converted into its trichloroacetimidate esters by reaction with allylic alcohols in  $\text{CH}_2\text{Cl}_2$ , in the presence of DBU at  $\sim 0^\circ$  to ambient temperatures [Anderson & Overman *J Am Chem Soc* **125** 12412 2003, Kirsch et al. *Org Lett* **9** 911 2007]; similarly prepared trichloroacetimidates of allylic alcohols were shown to undergo ether-directed Pd(II)-catalysed aza-Claisen rearrangements [Jamieson & Sutherland *Tetrahedron* **63** 2132 2007], and bis-trichloroacetimidates from 2-aminopropane-1,3-diols yielded dihydrooxazines through an acid catalysed cyclisation [Rondot et al. *Org Lett* **9** 247 2007].

**1,1,1-Trichloroethane** [71-55-6] **M 133.4, f  $-32.7^\circ$ , b  $74.0^\circ$ ,  $d_4^{20}$  1.337,  $n_D^{20}$  1.4385.** Wash it successively with conc HCl (or conc  $\text{H}_2\text{SO}_4$ ), aqueous 10%  $\text{K}_2\text{CO}_3$  ( $\text{Na}_2\text{CO}_3$ ), aqueous 10% NaCl, dry it with  $\text{CaCl}_2$  or  $\text{Na}_2\text{SO}_4$ , and fractionally distil it. It can contain up to 3% dioxane as preservative. This is removed by washing successively with 10% aqueous HCl, 10% aqueous  $\text{NaHCO}_3$  and 10% aqueous NaCl, and distilling over  $\text{CaCl}_2$  before use. [*Beilstein* **1** IV 138.]

**1,1,2-Trichloroethane** [79-00-5] **M 133.4, f  $-36.3^\circ$ , b  $113.6^\circ$ ,  $d_4^{20}$  1.435,  $n_D^{20}$  1.472.** Purify the chloroethane as for 1,1,1-trichloroethane above. [*Beilstein* **1** IV 139.]

**Trichloroethylene** [79-01-6] **M 131.4, f  $-88^\circ$ , b  $87.2^\circ$ ,  $d_4^{20}$  1.463,  $n_D^{21}$  1.4767.** Trichloroethylene undergoes decomposition in a similar way as  $\text{CHCl}_3$ , giving HCl, CO,  $\text{COCl}_2$  and organic products. It reacts with KOH, NaOH and 90%  $\text{H}_2\text{SO}_4$ , and forms azeotropes with water, MeOH, EtOH, and acetic acid. It is purified by washing successively with 2M HCl, water and 2M  $\text{K}_2\text{CO}_3$ , then dried with  $\text{K}_2\text{CO}_3$  and  $\text{CaCl}_2$ , then fractionally distilled before use. It has also been steam distilled from 10%  $\text{Ca}(\text{OH})_2$  slurry, most of the water being removed from the distillate by cooling to  $-30$  to  $-50^\circ$  and filtering off the ice through chamois skin: the trichloroethylene is then fractionally distilled at 250mm pressure and collected in a blackened container. [Carlisle & Levine *Ind Eng Chem (Anal Ed)* **24** 1164 1932, *Beilstein* **1** IV 712.]

**1,1,2-Trichloro-1,2,2-trifluoroethane** [76-13-1] **M 187.4, b  $47.6^\circ/760\text{mm}$ ,  $d_4^{20}$  1.576,  $n_D^{20}$  1.360.** Wash it with water, then with weak alkali. Dry it with  $\text{CaCl}_2$  or  $\text{H}_2\text{SO}_4$  and distil it. [Locke et al. *J Am Chem Soc* **56** 1726 1934, *Beilstein* **1** III 157, **1** IV 142.]

**Tridecanoic acid** [638-53-9] **M 214.4, m  $41.8^\circ$ ,  $44.5\text{--}45.5^\circ$  (several forms), b  $199\text{--}200^\circ/24\text{mm}$ ,  $\text{pK}_{\text{Est}} \sim 5.0$ .** Recrystallise the acid from acetone. [*Beilstein* **2** IV 1117.]

**7-Tridecanone (dihexyl ketone)** [462-18-0] **M 198.4, m  $33^\circ$ , b  $255^\circ/766\text{mm}$ .** Crystallise the ketone from EtOH. [*Beilstein* **1** H 715.]

**Tri-*n*-dodecylamine (Hydrogen ionophore I)** [102-87-4] **M 522.0, m  $15.7^\circ$ , b  $220\text{--}228^\circ/0.03\text{mm}$ ,  $d_4^{20}$  0.833,  $n_D^{20}$  1.4577,  $\text{pK}_{\text{Est}} \sim 11.0$ .** Distil tridodecylamine under high vacuum and  $\text{N}_2$ , and store it in the absence of  $\text{CO}_2$ . It can be recrystallised from 95%EtOH/\* $\text{C}_6\text{H}_6$  at low temperature under vacuum. The *hydrochloride* has m  $78\text{--}79^\circ$ . [Ra et al. *J Org Chem* **9** 259 1944, *Beilstein* **4** III 413, **4** IV 801.]

**Tri-*n*-dodecylammonium nitrate** [2305-34-2] **M 585.0.** Crystallise the salt from *n*-hexane/acetone (95:5) and keep it in a desiccator over  $\text{P}_2\text{O}_5$  under vacuum. [*Beilstein* **4** IV 801 for tridodecylamine.]

**Tri-*n*-dodecylammonium perchlorate** [5838-82-4] **M 622.4.** Recrystallise the salt from *n*-hexane or acetone and keep it in a desiccator over  $\text{P}_2\text{O}_5$ . (**Potentially explosive.**)

**Triethanolamine** [102-71-6] **M 149.2, m  $20\text{--}22^\circ$ , b  $190\text{--}103^\circ/5\text{mm}$ ,  $206\text{--}207^\circ/15\text{mm}$ ,  $335.4^\circ/760\text{mm}$ ,  $d_4^{20}$  1.124,  $n_D^{20}$  1.485,  $\text{pK}^{25}$  7.92.** Shake the amine gently with Linde type 4A molecular sieves for 24 hours, filter and fractionate it under a vacuum, and preferably in the presence of  $\text{N}_2$ . Store it in dark stoppered bottles under  $\text{N}_2$  as it is *hygroscopic*, and turns brown in air and light. It has a strong ammoniacal odour (like diethanolamine). It is miscible with  $\text{H}_2\text{O}$ , MeOH and  $\text{Me}_2\text{CO}$  and its solubilities at  $25^\circ$  in *n*-heptane,  $\text{Et}_2\text{O}$  and \* $\text{C}_6\text{H}_6$  are 0.4%, 1.6% and 4.2%, respectively. [See **diethanolamine** above, *Beilstein* **4** IV 1524.]

**Triethanolamine hydrochloride** [637-39-8] **M 185.7, m  $177^\circ$ ,  $\text{pK}^{25}$  7.92 (free base).** Crystallise the salt

from EtOH. Dry it at 80°. [*Beilstein* 4 IV 1525.]

**1,1,2-Triethoxyethane** [4819-77-6] **M 162.2, b 74°/28mm, 164°/~760mm, 167.2°/760mm, d<sub>4</sub><sup>20</sup> 0.897, n<sub>D</sub><sup>20</sup> 1.401.** Dry it with Na<sub>2</sub>SO<sub>4</sub> and distil it. [McElvain & Walters *J Am Chem Soc* 64 1964 1942, *Beilstein* 1 H 818, 1 I 418, 1 III 3184, 1 IV 3958.]

**Triethylamine** [121-44-8] **M 101.2, b 89.4°, d<sub>4</sub><sup>20</sup> 0.7280, n<sub>D</sub><sup>20</sup> 1.4005, pK<sup>25</sup> 10.82.** Dry triethylamine with CaSO<sub>4</sub>, LiAlH<sub>4</sub>, Linde type 4A molecular sieves, CaH<sub>2</sub>, KOH, or K<sub>2</sub>CO<sub>3</sub>, then distil it, either alone or from BaO, sodium, P<sub>2</sub>O<sub>5</sub> or CaH<sub>2</sub>. It has also been distilled from zinc dust, under nitrogen. To remove traces of primary and secondary amines, triethylamine has been refluxed with acetic anhydride, benzoic anhydride, phthalic anhydride, then distilled, refluxed with CaH<sub>2</sub> (ammonia-free) or KOH (or dried with activated alumina), and again distilled. Another purification method involved refluxing for 2 hours with *p*-toluenesulfonyl chloride, then distilling. Grovenstein and Williams [*J Am Chem Soc* 83 412 1961] treated triethylamine (500ml) with benzoyl chloride (30ml), filtered off the precipitate, and refluxed the liquid for 1 hour with a further volume of benzoyl chloride (30ml). After cooling, the liquid was filtered, distilled, and allowed to stand for several hours with KOH pellets. It was then refluxed with, and distilled from, stirred molten potassium. Triethylamine has been converted to its *hydrochloride* (see below), crystallised from EtOH (to **m** 254°), then liberated with strong aqueous NaOH, dried with solid KOH and distilled from sodium under N<sub>2</sub>. It is a strong base and should not be inhaled. [*Beilstein* 4 H 99, 4 I 348, 4 II 593, 4 III 194, 4 IV 322.]

**Triethylammonium hydrobromide** [636-70-4] **M 229.1, m 248°.** Equimolar portions of triethylamine and aqueous solutions of HBr in acetone are mixed with cooling. The precipitated salt is washed with anhydrous acetone and dried in a vacuum for 1-2 hours. [Odinekov et al. *J Chem Soc, Faraday Trans 2* 80 899 1984.] Recrystallise it from CHCl<sub>3</sub> or EtOH. [*Beilstein* 4 IV 322.]

**Triethylammonium hydrochloride** [554-68-7] **M 137.7, m 257-260°(dec).** Purify it like the bromide above. [*Beilstein* 4 IV 327.]

**Triethylammonium hydroiodide** [4636-73-1] **M 229.1, m 181°.** Purify it as for triethylammonium bromide, except the solution for precipitation should be precooled acetone at -10°, and the precipitate is twice recrystallised from a cooled acetone/hexane mixture at -10°. Store it in the dark. [*Beilstein* 4 IV 327.]

**Triethylammonium trichloroacetate** [4113-06-8] **M 263.6.** Equimolar solutions of triethylamine and trichloroacetic acid in *n*-hexane are mixed at 10°. The solid so obtained is recrystallised from CHCl<sub>3</sub>/\*benzene. [Hoigbne & Gäumann *Helv Chim Acta* 42 444 1959, *Beilstein* 4 IV 330.]

**Triethylammonium trifluoroacetate** [454-49-9] **M 196.2.** Purify as for the corresponding trichloroacetate but in Et<sub>2</sub>O. Evaporation of the Et<sub>2</sub>O gives the salt as a colourless viscous liquid at ambient temperature. [Emmons et al. *J Am Chem Soc* 76 3472 1954, *Beilstein* 4 IV 330.]

**Triethylene glycol** [112-27-6] **M 150.2, b 115-117°/0.1mm, 278°/760mm, d<sub>4</sub><sup>15</sup> 1.1274, n<sub>D</sub><sup>15</sup> 1.4578.** Dry the glycol with CaSO<sub>4</sub> for 1 week, then it is repeatedly and very slowly fractionally distilled under a vacuum. Store it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. It is very *hygroscopic*. [*Beilstein* 1 IV 2400.]

**Triethylene glycol dimethyl ether (triglyme)** [112-49-2] **M 178.2, b 225°, d<sub>4</sub><sup>20</sup> 0.987, n<sub>D</sub><sup>20</sup> 1.425.** Reflux it with, and distil it from sodium hydride or LiAlH<sub>4</sub>. [*Beilstein* 1 IV 2401.]

**Triethylenetetramine (TRIEN, TETA, trientine, H<sub>2</sub>NCH<sub>2</sub>CH<sub>2</sub>NHCH<sub>2</sub>CH<sub>2</sub>NHCH<sub>2</sub>CH<sub>2</sub>NHCH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub>)** [112-24-3] **M 146.2, m 12°, b 157°/20mm, d<sub>4</sub><sup>20</sup> 0.971, n<sub>D</sub><sup>20</sup> 1.497, pK<sub>1</sub><sup>25</sup> 3.32, pK<sub>2</sub><sup>25</sup> 6.67, pK<sub>3</sub><sup>25</sup> 9.20, pK<sub>4</sub><sup>25</sup> 9.92.** Dry the amine with sodium, then distil it under a vacuum. Further purification has been *via* the nitrate or the chloride salts. For example, Jonassen and Strickland [*J Am Chem Soc* 80 312 1958] separated TRIEN from admixture with TREN (38%) by solution in EtOH, cooling to approximately 5° in an ice-bath and adding conc HCl dropwise from a burette, keeping the temperature below 10°, until all of the white crystalline precipitate of TREN.HCl (see p 216) had formed and was removed. Further addition of HCl then precipitated thick, creamy

white TRIEN.HCl (see below) which was crystallised several times from hot water by adding an excess of cold EtOH. The crystals were finally washed with Me<sub>2</sub>CO, then Et<sub>2</sub>O and dried in a vacuum desiccator. [Beilstein 4 H 255, 4 II 695, 4 III 542, 4 IV 1242.]

**Triethylenetetramine tetrahydrochloride (TRIEH HCl)** [4961-10-4] **M 292.1, m 266-270°**. Crystallise the salt repeatedly from hot water by precipitation with cold EtOH or EtOH/HCl. Wash it with acetone and absolute EtOH and dry it in a vacuum oven at 80° (see TRIEN above). The *tetrabenzoyl* derivative has **m 230°** after recrystallisation from boiling EtOH. [Peacock *J Chem Soc* 1519 1936, *Beilstein* 4 H 255, 4 II 695, 4 III 543.]

**Triethyl orthoformate (ethyl orthoformate, 1,1,1-triethoxymethane)** [122-51-0] **M 148.2, m 30°, b 60°/30mm, 144-146°/760mm, d<sub>4</sub><sup>20</sup> 0.891, n<sub>D</sub><sup>20</sup> 1.392**. Fractionate it first at atmospheric pressure, then in a vacuum. If impure, then shake it with aqueous 2% NaOH, dry it with solid KOH and distil it from sodium through a 20cm Vigreux column. *Alternatively*, wash it with H<sub>2</sub>O, dry it over anhydrous K<sub>2</sub>CO<sub>3</sub>, filter and fractionate it through a Widmer column. [Sah & Ma *J Am Chem Soc* 54 2964 1932, Ohme & Schmitz *Justus Liebigs Ann Chem* 716 207 1968, *Beilstein* 2 IV 25.] **IRRITANT** and **FLAMMABLE**.

**Triethyloxonium fluoroborate** [368-39-8] **M 190.0, m 92-93°(dec)**. Crystallise it from diethyl ether. It is *very hygroscopic*, and must be handled in a dry box and stored at 0°. [Meerwein *Org Synth Coll Vol V* 1096 1973.] Pure material should give a clear and colourless solution in dichloromethane (1 in 50, w/v). [Beilstein 1 IV 1322.]

**Trifluoroacetic acid** [76-05-1] **M 114.0, f -15.5°, b 72.4°, d<sub>4</sub><sup>20</sup> 1.494, n<sub>D</sub><sup>20</sup> 1.2850, pK<sup>25</sup> 0.52**. Although an improved preparation of the acid (in 87% yield) involving the oxidation of 2,3-dichloro-1,1,1,4,4,4-hexafluorobut-2-ene (b 65-66°/745mm, [303-04-8], steam volatile, **d<sub>4</sub><sup>25</sup> 1.605, n<sub>D</sub><sup>20</sup> 1.3458**) *via* oxidation with KMnO<sub>4</sub>/KOH/95°/8-10 hours has been reported [Henne & Trott *J Am Chem Soc* 69 1820 1947], the next purification procedure should be avoided. The purification of trifluoroacetic acid, reported in earlier editions of this work, by refluxing over KMnO<sub>4</sub> for 24 hours and slowly distilling has resulted in very **SERIOUS EXPLOSIONS** on various occasions, but not always. This apparently depends on the source and/or age of the acid. The method is **NOT RECOMMENDED**. It forms an azeotropic mixture of acid and H<sub>2</sub>O (80:20) which boils at 103-105°/745mm. Water can be removed by adding trifluoroacetic anhydride (0.05%, to diminish water content) and distilling. [Conway & Novak *J Phys Chem* 81 1459 1977]. It can be refluxed and distilled from P<sub>2</sub>O<sub>5</sub>, but do not use an excess of P<sub>2</sub>O<sub>5</sub> as it produces the anhydride (see below and Schmidt & Staub *Chem Ber* 87 388 1954) which can be separated by fractional distillation because the anhydride distils about 30° lower at atmospheric pressure. It is further purified by fractional crystallisation by partial freezing and again distilled. It is a strong acid and attacks skin (wear gloves) and eyes (wear safety glasses). **Highly TOXIC vapour — do not inhale it. Work in an efficient fume hood.** [Beilstein 2 II 186, 2 III 426, 2 IV 458.]

**Trifluoroacetic anhydride** [407-25-0] **M 210.0, b 38-40°/760mm, d<sub>4</sub><sup>20</sup> 1.508**. Purification by distilling over KMnO<sub>4</sub>, as for the acid above, is **EXTREMELY DANGEROUS** due to the possibility of **EXPLOSION** (see preceding acid). It is best purified by distilling from P<sub>2</sub>O<sub>5</sub> slowly, and collecting the fraction boiling at 39.5°. Store it in a dry atmosphere. **Highly TOXIC vapour and attacks skin, work in an efficient fume hood — see previous entry.** [Tedder *Chem Rev* 55 787, *Beilstein* 2 II 186, 2 III 427, 2 IV 469.]

**1,1,1-Trifluoroacetone** [421-50-1] **M 112.1, b 21°/atm, 22°/atm, d<sub>4</sub><sup>25</sup> 1.252, n<sub>D</sub><sup>20</sup> 1.3**. The ketone was obtained by refluxing ethyl 4,4,4-trifluoroacetoacetate (see [372-31-6]) for 2 hours with 10% aqueous H<sub>2</sub>SO<sub>4</sub> and collecting the distillate into a receiver containing P<sub>2</sub>O<sub>5</sub>, and redistilling it from P<sub>2</sub>O<sub>5</sub>. The *2,4-dinitrophenylhydrazone* has **m 139°** [Henne et al. *J Am Chem Soc* 69 1819 1947]. *Alternatively*, ethyl 4,4,4-trifluoro-sodio-acetoacetate, prepared from the ester with one equivalent of NaOEt was heated with hot (80-90°) dilute H<sub>2</sub>SO<sub>4</sub> in an all-glass apparatus to give trifluoroacetone which was condensed in dry Et<sub>2</sub>O at -78°. The ethereal solution was dried (MgSO<sub>4</sub> or CaCl<sub>2</sub>), distilled, and the desired product condensed at -78° [Haszeldine et al. *J Chem Soc* 609 1951]. The vapours are very **TOXIC (work in an efficient fume cupboard)**. The *semicarbazone* has, **m 129-130°**, and the *2,4-dinitrophenylhydrazone* has **m 141-142°**. [McBee et al. *J Am Chem*

*Soc* 75 4090 1953, *Beilstein* 1 II 717, 1 III 2745, 1 IV 3214.] Trifluoroacetone has been used to prepare trifluoromethylamines of chiral amines in order to prepare *R* or *S* enantiopure  $\alpha$ -trifluoromethyl alanines, diamines and amino alcohols via a Strecker-type synthesis in a few steps [Huguenot & Brigaud *J Org Chem* 71 7075 2006]. It has also been used to synthesise 2-trifluoromethyl-7-azaindoles from 2,6-dihalopyridines [Schirok et al. *Synthesis* 251 2007].

**2,2,2-Trifluoroethanol** [75-89-8] **M 100.0, m  $-44^\circ$ , b  $72.4^\circ/738\text{mm}$ ,  $77-80^\circ/\text{atm}$ ,  $n_D^{20}$  1.30,  $d_4^{20}$  1.400,  $pK^{25}$  12.8.** Dry it with  $\text{CaSO}_4$  and a little  $\text{NaHCO}_3$  (to remove traces of acid) and distil it. It is used as a calibration standard for NMR spectroscopy. **Highly TOXIC vapour.** [*Beilstein* 1 IV 1370.]

**Trifluoromethanesulfonic anhydride (triflic anhydride)** [358-23-6] **M 282.1, b  $82-85^\circ$ ,  $84^\circ$ ,  $d_4^{20}$  1.71,  $n_D^{20}$  1.322.** Distil it through a short Vigreux column. It can be freshly prepared from the anhydrous acid (11.5g) and  $\text{P}_2\text{O}_5$  (11.5g, or half this weight) by setting aside at room temperature for 1 hour, distilling off volatile products then distil it through a short Vigreux column. It is readily hydrolysed by  $\text{H}_2\text{O}$  and decomposes appreciably after a few days to liberate  $\text{SO}_2$  and produce a viscous liquid. Store it dry at low temperatures. [Burdon et al. *J Chem Soc* 2574 1957, Beard et al. *J Org Chem* 38 373 1973, *Beilstein* 3 IV 35.] **Highly TOXIC vapour.**

**1,1,1-Trifluoro-2,4-pentanedione ( $\alpha,\alpha,\alpha$ -trifluoroacetylacetone, *tfacac* or *facac*)** [367-57-7] **M 154.1, b  $105-107^\circ/\text{atm}$ ,  $107^\circ/760\text{mm}$ ,  $d_4^{25}$  1.27,  $n_D^{20}$  1.3890.** The diketone was prepared from ethyl trifluoroacetate (74.5g) and Na wire (10g, or an equivalent of  $\text{NaOEt}$ ) in  $\text{Et}_2\text{O}$  (5ml) at  $0^\circ$ , by adding  $\text{Me}_2\text{CO}$  (80ml) at a rate to maintain a steady reaction, kept at  $0^\circ$  for 2 hours, and then allowed to warm to room temperature. Ice-water (5ml) was added, the aqueous layer was acidified with  $\text{AcOH}$ , and the Cu complex was precipitated with a warm saturated solution of  $\text{Cu}(\text{OAc})_2$ . The *copper trifluoroacetylacetone* crystallised in long blue needles from  $\text{EtOH}$  and had **m  $189^\circ$** ; its UV has  $\lambda_{\text{max}}$  at 240nm ( $\epsilon$  19,500) and 294nm ( $\epsilon$  31,000) in  $\text{EtOH}$ , and the effect of fluorine substitution on the complexing properties has been studied in detail by Calvin and coworkers [Belford et al. *J Inorg Nucl Chem* 2 11 1956]. An ethereal mixture of the complex (partly dissolved) was shaken with 2N  $\text{H}_2\text{SO}_4$ , then  $\text{H}_2\text{O}$ , the organic layer was separated, dried ( $\text{Na}_2\text{SO}_4$ ), filtered and distilled to give a 62% yield of *trifluoroacetylacetone*. Alternatively, the metal can be removed by bubbling  $\text{H}_2\text{S}$  through the ethereal mixture, the  $\text{CuS}$  was filtered off and the filtrate was distilled. [Reid & Calvin *J Am Chem Soc* 72 2948 1950, Henne et al. *J Am Chem Soc* 69 1819 1947, Haszeldine et al. *J Chem Soc* 609 1951, Belford et al. *J Inorg Nucl Chem* 2 11 1956]. Its UV has  $\lambda_{\text{max}}$  at 284nm ( $\epsilon$  10,500) in  $\text{CHCl}_3$ ; the FT-IR (film) has  $\nu_{\text{max}}$  at 1604.7, 1422.3, 1369.5, 1282.2, 1154.7, 887.4, 857.4, 795.2, 726.4  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 2.36 (s), 4.60 (s, 3H,  $\text{CH}_3$ ), 8.30 (s, CH); and the  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 193.47, 176.71, 176.22, 175.74, 175.25, 127.75, 119.00, 115.25, 111.50, 96.37, 96.35, 24.83. It enolises in aqueous solution, and rates of enolisation were determined by Ried and Calvin (*vide infra*) in  $\text{H}_2\text{O}$ , 0.1N and 0.5N  $\text{HCl}$ , and in aqueous  $\text{EtOH}$  of varying concentrations; and are temperature dependent. Thus at equilibrium, at 0.08-0.09M solution of *tfacac* and  $25^\circ$ , 16% aqueous  $\text{EtOH}$  contains 2.5% of enol+enolate, whereas at 95% it contains 20.7% of enol+enolate. This enolisation should be taken into account when measuring spectra and other physical parameters. Complexation with Cu, Fe and Al can be used to extract the metals from aqueous medium (acetate buffered) by a 0.10M of the diketone in  $\text{CHCl}_3$  [Scribner et al. *Anal Chem* 37 1136 1965], and Be into organic solvents [Scribner et al. *Anal Chem* 38 1779 1966]. The *oxime*, **m  $86-87^\circ$** , crystallises from  $\text{H}_2\text{O}$  or aqueous  $\text{EtOH}$ . [*Beilstein* 1 III 3123, 1 IV 3680.]

**Trimethylamine** [75-50-3] **M 59.1, b  $3.5^\circ$ ,  $pK^{25}$  9.80.** Dry triethylamine by passing the gas through a tower filled with solid  $\text{KOH}$ . Water and impurities containing labile hydrogen were removed by treatment with freshly sublimed, ground,  $\text{P}_2\text{O}_5$ . It has been refluxed with acetic anhydride, and then distilled through a tube packed with  $\text{HgO}$  and  $\text{BaO}$ . [Comyns *J Chem Soc* 1557 1955.] For more extensive purification, trimethylamine is converted to the hydrochloride, crystallised (see below), and regenerated by treating the hydrochloride with excess aqueous 50%  $\text{KOH}$ ; the gas is passed through a  $\text{CaSO}_4$  column into a steel cylinder containing sodium ribbon. After 1-2 days, the cylinder is cooled to  $-78^\circ$  and hydrogen and air are removed by pumping. [Day & Felsing *J Am Chem Soc* 72 1698 1950.]  $\text{Me}_3\text{N}$  has been distilled from trap-to-trap methods and degassed by freeze-pump-thaw [Halpern et al. *J Am Chem Soc* 108 3907 1986]. It is commercially available under pressure

in tin cylinders. [*Beilstein* 4 H 43, 4 I 322, 4 II 553, 4 III 99, 4 IV 134.]

**Trimethylamine hydrochloride** [593-81-7] **M 95.7, m >280°(dec)**. The salt crystallises from CHCl<sub>3</sub>, EtOH or *n*-propanol, and is dried under vacuum. It also crystallises from \*benzene/MeOH, MeOH/diethyl ether and is dried under vacuum over paraffin wax and H<sub>2</sub>SO<sub>4</sub>. It is kept over P<sub>2</sub>O<sub>5</sub> as it is *hygroscopic*. [*Beilstein* 4 H 262, 4 I 419, 4 IV 138.]

**Trimethylamine hydroiodide** [20230-89-1] **M 186.0, m 263°**. It crystallises from MeOH.

**Trimethylolpropane (1,1,1-tris-hydroxymethylpropane, 2-ethyl-2-hydroxymethyl-1,3-propanediol)** [77-99-6] **M 134.2, m 57-59°, 60-62°, b 159-161°/2mm**. Crystallise it from acetone and ether, and it distils at high vacuum. [*Beilstein* 1 III 2349.]

**2,2,3-Trimethylpentane** [564-02-3] **M 114.2, b 109.8°, d<sub>4</sub><sup>20</sup> 0.7161, n<sub>D</sub><sup>20</sup> 1.40295, n<sub>D</sub><sup>25</sup> 1.40064**. Purify it by azeotropic distillation with 2-methoxyethanol, which is subsequently washed out with water. The trimethylpentane is then dried and fractionally distilled. [Forziati et al. *J Res Nat Bur Stand* 36 129 1946, *Beilstein* 1 IV 439.]

**2,2,4-Trimethylpentane (isooctane)** [540-84-1] **M 114.2, m -107°, b 99.2°, d<sub>4</sub><sup>20</sup> 0.693, n<sub>D</sub><sup>20</sup> 1.39145, n<sub>D</sub><sup>25</sup> 1.38898**. Distil isooctane from sodium, pass it through a column of silica gel or activated alumina (to remove traces of olefins), and again distilled from sodium. Extract it repeatedly with conc H<sub>2</sub>SO<sub>4</sub>, then agitate it with aqueous KMnO<sub>4</sub>, wash it with water, dry (CaSO<sub>4</sub>) and distil it. Purify it also by azeotropic distillation with EtOH, which is subsequently washed out with water, and the trimethylpentane is dried and fractionally distilled. [Forziati et al. *J Res Nat Bur Stand* 36 126 1946.] [*Beilstein* 1 IV 439.]

**2,4,4-Trimethylpent-2-ene (β-diisobutylene)** [107-40-4] **M 112.2, m -106°, b 104°, d<sub>4</sub><sup>20</sup> 0.720, n<sub>D</sub><sup>20</sup> 1.4160**. Fractionate it under N<sub>2</sub> as it is highly flammable. [*Beilstein* 1 III 848, 1 IV 891.]

**Trimethylsulfonium iodide** [2181-42-2] **M 204.1, m 211-212.5°(dec), 215-220°(dec)**. Crystallise the iodide from EtOH. [Emeleus & Heal *J Chem Soc* 1126 1946, Swain & Kaiser *J Am Chem Soc* 80 4089 1958, Borredon et al. *J Org Chem* 55 501 1990, Bouda et al. *Synth Commun* 17 503 1987.]

**Trimyristin** [555-45-3] **M 723.2, m 56.5°**. Crystallise it from diethyl ether. [*Beilstein* 2 IV 1135.]

**Tri-*n*-octylamine** [1116-76-3] **M 353.7, b 164-168°/0.7mm, 365-367°/760mm, d<sub>4</sub><sup>20</sup> 0.813, n<sub>D</sub><sup>20</sup> 1.450, pK<sup>25</sup> 10.65**. It is converted to the amine hydrochloride etherate which is recrystallised four times from diethyl ether at -30° (see below). Neutralisation of this salt regenerates the free amine which distil under high vacuum. [Wilson & Wogman *J Phys Chem* 66 1552 1962.] Distil the strong base amine at <1-2mm pressure. [*Beilstein* 4 H 196, 4 III 382, 4 IV 754.]

**Tri-*n*-octylammonium chloride** [1188-95-0] **M 384.2, m 78-79°, pK<sup>25</sup> 8.35 (in 70% aqueous EtOH)**. Crystallise it from Et<sub>2</sub>O, then *n*-hexane (see above). [Burrows et al. *J Chem Soc* 200 1947, *Beilstein* 4 H 196.]

**Tri-*n*-octylammonium perchlorate** [2861-99-6] **M 454.2, m >300°(dec)**. Crystallise the perchlorate from *n*-hexane. (Possibly explosive.) [*Beilstein* 4 IV 754.]

**Tripalmitin** [555-44-2] **M 807.4, m 66.4°**. Crystallise it from acetone, diethyl ether or EtOH. It exists in an α-form (m 56.0°), a β'-form (m 63.5°) and a β-form (m 65.5°). [*Beilstein* 2 H 373, 2 I 167, 2 II 340, 2 III 971.]

**Tri-*n*-propylamine** [102-69-2] **M 143.3, b 156.5°, d<sub>4</sub><sup>20</sup> 0.757, n<sub>D</sub><sup>20</sup> 1.419, pK<sup>25</sup> 10.66**. Dry the amine with KOH and fractionally distil it. Also reflux it with toluene-*p*-sulfonyl chloride and with KOH, then fractionally distil it. The distillate, after addition of 2% phenyl isocyanate, was redistilled and the residue fractionally distilled from sodium. It is a strong base. [Takahashi et al. *J Org Chem* 52 2666 1987, *Beilstein* 4 IV 470.]

**Tris(2-aminoethyl)amine (TREN)** [4097-89-6] **M 146.2, b 114°/15mm, 263°/744mm, d<sub>4</sub><sup>20</sup> 0.977, n<sub>D</sub><sup>20</sup> 1.498, pK<sub>1</sub><sup>25</sup> 8.42, pK<sub>2</sub><sup>25</sup> 9.44, pK<sub>3</sub><sup>25</sup> 10.13.** For a separation from a mixture containing 62% TRIEN, see entry under triethylenetetramine [112-24-3] above. Also purify it by conversion to the hydrochloride (see below), recrystallise it and regenerate the free base [Xie & Hendrickson *J Am Chem Soc* **109** 6981 1987]. [*Beilstein* **4 H** 256, **4 II** 695, **4 III** 545, **4 IV** 1250.]

**Tris(2-aminoethyl)amine trihydrochloride (TREN.HCl)** [14350-52-8] **M 255.7, m 300°(dec).** Crystallise the salt several times by dissolving it in the minimum of hot water and precipitating it with excess of cold EtOH. The precipitate is washed with acetone, then diethyl ether and dried in a vacuum desiccator. [*Beilstein* **4 H** 256, **4 II** 695, **4 III** 545, **4 IV** 1250.]

**Tris(dimethylamino)methane (N,N,N',N',N'',N''-hexamethylmethanetriamine)** [5762-56-1] **M 145.3, b 42-43°/12mm, n<sub>D</sub><sup>20</sup> 1.4349, pK<sub>Est</sub> ~10.** Dry it over KOH and distil it through a Vigreux column at water pump vacuum. Store it in the absence of CO<sub>2</sub>. [Bredereck et al. *Chem Ber* **101** 1885 1968 and *Angew Chem, Int Ed Engl* **5** 132 1966.]

**Tris(hydroxymethyl)methylamine (TRIS)** [77-86-1] **M 121.1, m 172°, pK<sup>25</sup> 8.07.** TRIS can ordinarily be obtained in highly pure form suitable sources for use as an acidimetric standard. If only impure material is available, it should be crystallised from 20% EtOH, aqueous MeOH (**m** 171.1°) or isopropanol (**m** 172-173°). Dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or CaCl<sub>2</sub>. *Alternatively*, it is dissolved in twice its weight of water at 55-60°, filtered, concentrated to half its volume and poured slowly, with stirring, into about twice its volume of EtOH. The crystals which separate on cooling to 3-4° are filtered off, washed with a little MeOH, air dried by suction, then finally ground and dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. It has also been recrystallised from water, MeOH or aqueous MeOH, and vacuum dried at 80° for 2 days. [*Beilstein* **4 H** 303, **4 III** 857, **4 IV** 1903.]

**Tris(hydroxymethyl)methylammonium hydrochloride (TRIS-HCl)** [1185-53-1] **M 157.6, m 149-150°(dec).** Crystallise the salt from 50% EtOH, then from 70% EtOH. TRIS-hydrochloride is also available commercially in a highly pure state. Otherwise, recrystallise it from 50% EtOH, then 70% EtOH, and dry it below 40° to avoid risk of decomposition. [*Beilstein* **4 H** 304.]

**1,1,1-Tris(hydroxymethyl)ethane (2-hydroxymethyl-2-methyl-1,3-propanediol)** [77-85-0] **M 120.2, m 200°.** Dissolve it in hot tetrahydrofuran, filter and precipitate it with hexane. It has also been crystallised from acetone/water (1:1). Dry it in a vacuum. [*Beilstein* **1 H** 520, **1 IV** 2780.]

**N-Tris(hydroxymethyl)methyl-2-aminomethanesulfonic acid (TES)** [7365-44-8] **M 229.3, m 224-226°(dec), pK<sup>20</sup> 7.50.** Crystallise the acid from hot EtOH containing a little water.

**Tris(hydroxymethyl)nitromethane [2-(hydroxymethyl)-2-nitro-1,3-propanediol]** [126-11-4] **M 151.1, m 174-175°(dec, tech. grade), 214°(pure).** Crystallise it from CHCl<sub>3</sub>/ethyl acetate or ethyl acetate/\*benzene. It is an acid and a 0.1M solution in H<sub>2</sub>O has pH 4.5. **IRRITANT.** [*Beilstein* **1 H** 520.]

**Tris[2-(methylamino)ethyl]amine** [65604-89-9] **M 188.3, b 77-78°/0.1mm, d<sub>4</sub><sup>20</sup> 0.896, pK<sub>Est(1)</sub> ~8.8, pK<sub>Est(2)</sub> ~9.4, pK<sub>Est(3)</sub> ~10.4.** If this strong base contains carbonate (check IR) it should be shaken with solid KOH, decanted and distilled at high vacuum to give a colourless oil with a strong amine odour. Store it in the dark under N<sub>2</sub> as it absorbs CO<sub>2</sub> in moist air. It is synthesised in two steps. Ethyl chloroformate (33.4g, 310mmol, lachrymatory, see [541-41-3] above) is added dropwise to a mixture of TREN (29.2g, 200mmol, see [4097-89-6] above) in \*C<sub>6</sub>H<sub>6</sub> (225ml) and H<sub>2</sub>O (100ml), and cooled to 5°. Then a solution of KOH (36.4g, 650mmol) in H<sub>2</sub>O (35ml) is added simultaneously with more ethyl chloroformate (33.4g, 310mmol), with stirring, while keeping the reaction mixture below 5° for 2 hours followed, still with stirring, by 8 hours at ~25°. The \*C<sub>6</sub>H<sub>6</sub> layer is separated, the aqueous layer is extracted with CHCl<sub>3</sub> (3 x 100ml), the combined organic layers are dried (MgSO<sub>4</sub>), filtered and the filtrate is evaporated to leave *tris[2-(ethoxycarbonylamino)ethyl]amine* (~85%) as a thick oil which is used directly in the subsequent step. [The crude tri-ester has IR (film) bands with  $\nu_{\max}$  at

3300, 1720, 1530, and 1250  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.27 (t, 9H,  $^3J_{\text{HH}} = 7.1\text{Hz}$ ), 2.60 (t, 6H,  $^3J_{\text{HH}} = 5.7\text{Hz}$ ), 3.23 (br s, 6H), 4.10 (q, 6H,  $^3J_{\text{HH}} = 7.1\text{Hz}$ ) and 5.50 (br s, 3H)].

In the second step the preceding crude tri-ester (61.3g, 170mmol) in THF (250ml) is added dropwise to a suspension of  $\text{LiAlH}_4$  (30.0g, 790mmol) in THF (700ml), and the reaction mixture is refluxed overnight. Water (50ml), followed by a solution of KOH (50g) in  $\text{H}_2\text{O}$  (50ml) are very carefully added to it (cool if necessary), the solvent is decanted from the inorganic gel, evaporated *in vacuo*, and the residual yellow oil is fractionated in a vacuum to give the desired *amine* in 88% yield. It has  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ , TMS) with  $\delta$  at 1.30 (br s, 3H, NH), 2.39 (s, 9H,  $\text{CH}_3$ ), 2.48 (m, 6H,  $^3J_{\text{HH}} = 6.1\text{Hz}$ , 3 $\text{CH}_2$ ) and 2.52 (m, 6H,  $^3J_{\text{HH}} = 6.1\text{Hz}$ , 3 $\text{CH}_2$ ); the  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 54.1 ( $\text{CH}_2$ ), 49.6 ( $\text{CH}_2$ ) and 36.3 ( $\text{CH}_3$ ); and HRMS has  $m/z$  189.2082 (calc for M + H is 189.20793). [Schmidt et al. *Z Anorg Allg Chem* **578** 75 1989.]

**Triuret (1,3-dicarbamoylurea)** [556-99-0] **M 146.1, m 233 $^\circ$ (dec)**. It crystallises from aqueous ammonia or  $\text{H}_2\text{O}$  (plates **m 232-234 $^\circ$** ). It gives mono and dipotassium salts. [Beilstein **3** H 72, **3** I 35, **3** II 60, **3** III 142.]

**Undecan-1-ol** [112-42-5] **M 172.3, m 16.5 $^\circ$ , 146 $^\circ$ /30mm,  $d_4^{25}$  0.830,  $n_D^{20}$  1.440**. Purify the alcohol by repeated fractional crystallisation from its melt or by distillation in a vacuum. [Beilstein **1** H 427, **1** IV 1835.]

**Undecanoic acid (C11, undecylic acid)** [112-78-8] **M 186.3, m 28.5 $^\circ$ , b 164 $^\circ$ /18mm, 228 $^\circ$ /160mm, 248-250 $^\circ$ /~760mm,  $d_4^{20}$  0.8907,  $n_D^{25}$  1.4294,  $\text{pK}_{\text{Est}} \sim 5.0$** . Purify the acid by repeated fractional crystallisation from its melt or by distillation in a vacuum. [Beilstein **2** H 358, **2** IV 1068.]

**Undec-10-enoic acid** [112-38-9] **M 184.3, m 25-25.5 $^\circ$ , b 131 $^\circ$ /1mm, 168 $^\circ$ /15mm,  $d_4^{20}$  0.912,  $n_D^{25}$  1.447,  $\text{pK}_{\text{Est}} \sim 5.0$** . Purify the acid by repeated fractional crystallisation from its melt or by distillation, preferably in a high vacuum. [Beilstein **2** IV 1612.]

**Urea** [57-13-6] **M 60.1, m 132.7-132.9 $^\circ$ ,  $\text{pK}^{25}$  0.12**. Crystallise urea twice from conductivity water using centrifugal drainage and keeping the temperature below 60 $^\circ$ . The crystals are dried under vacuum at 55 $^\circ$  for 6 hours. Levy and Margouls [*J Am Chem Soc* **84** 1345 1962] prepared a 9M solution in conductivity water (keeping the temperature below 25 $^\circ$ ) and, after filtering through a medium-porosity glass sinter, added an equal volume of absolute EtOH. The mixture was set aside at -27 $^\circ$  for 2-3 days and filtered cold. The precipitate was washed with a small amount of EtOH and dried in air. Crystallisation from 70% EtOH between 40 $^\circ$  and -9 $^\circ$  has also been used. Ionic impurities such as ammonium isocyanate have been removed by treating the concentrated aqueous solution at 50 $^\circ$  with Amberlite MB-1 cation- and anion-exchange resin, and allowing it to crystallise on evaporation. [Benesch et al. *J Biol Chem* **216** 663 1955.] It can also be crystallised from MeOH or EtOH, and is dried under vacuum at room temperature. [Beilstein **3** H 42, **3** I 19, **3** II 35, **3** III 80.]

**Urea nitrate** [124-47-0] **M 123.1, m 152 $^\circ$ (dec), 157-158 $^\circ$ , 163 $^\circ$** . Crystallise it from dilute  $\text{HNO}_3$  or EtOH (**m 157-158 $^\circ$** ), and dry it in a vacuum over  $\text{P}_2\text{O}_5$ . [Beilstein **3** H 54, **3** I 25, **3** II 45, **3** III 105, **3** IV 94.]

**Urethane (ethyl carbamate, ethyl urethane)** [51-79-6] **M 89.1, m 48-50 $^\circ$ , b 182-184 $^\circ$ /~760mm,  $d_4^{20}$  0.986,  $n_D^{25}$  1.4144**. Urethane is best purified by fractional distillation, but it can be sublimed at ~103 $^\circ$ /~50mm. It has also been recrystallised from \*benzene. Its solubility at room temperature is 2g/ml in  $\text{H}_2\text{O}$ , 1.25g/ml in EtOH, 1.1g/ml in  $\text{CHCl}_3$ , 0.67g/ml in  $\text{Et}_2\text{O}$  and 0.03g/ml in olive oil. It is a suspected **human carcinogen**. [Beilstein **3** H 22, **3** IV 40.]

**cis-Vaccenic acid (octadec-11-enoic acid)** [506-17-2] **M 282.5, m 14-15 $^\circ$ , b 158-163 $^\circ$ /0.4mm,  $d_4^{20}$  0.880,  $n_D^{25}$  1.4598,  $\text{pK}_{\text{Est}} \sim 4.9$** . Purify the acid by fractional distillation under high vacuum or crystallisation from its melt in an inert atmosphere away from light. [Beilstein **2** I 198, **2** III 1384, **2** IV 1639.]

**trans-Vaccenic acid (octadec-11-enoic acid)** [693-72-1] **M 282.5, m 43-44 $^\circ$ ,  $n_D^{50}$  1.4472,  $\text{pK}_{\text{Est}} \sim 4.9$** . Crystallise the acid from acetone (**m 45-45.5 $^\circ$** ) or aqueous MeOH (**m 43.5-43.7 $^\circ$** ). The *methyl ester* has **b 174-**

175°/5mm. [Böeseken & Hoagland *Rec Trav Chim Pays Bas* **46** 632 1927, Ahmad et al. *J Am Chem Soc* **70** 3391 1948, IR: Rao & Daubert *J Am Chem Soc* **70** 1102 1948.]

***n*-Valeraldehyde (pentanal)** [110-62-3] **M 86.1, m -92°, b 103°, d<sub>4</sub><sup>20</sup> 0.811, n<sub>D</sub><sup>25</sup> 1.40233**. Purify pentanal via the bisulfite derivative (see **2-butanone** above for the preparation and decomposition of the bisulfite derivative). [Birrell & Trotman-Dickinson *J Chem Soc* 2059 1960, *Beilstein* **1** H 676, **1** IV 3268.] The **2,4-dinitrophenylhydrazone** [2057-84-3] **M 266.3** has **m 103-105°** (from EtOH). [*Beilstein* **15** III/IV 429.]

***n*-Valeramide (pentanamide)** [626-97-1] **M 101.1, m 115-116°**. Crystallise the amide from EtOH. It sublimes at 80°. [Philbrook *J Org Chem* **19** 624 1954, *Beilstein* **2** H 301, **2** I 131, **2** II 266, **2** III 674, **2** IV 874.]

**Valeric acid (*n*-pentanoic acid)** [109-52-4] **M 102.1, b 95°/22mm, 186.4°/~760mm, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>20</sup> 1.4080, pK<sup>25</sup> 4.81**. Water is removed from the acid by distillation using a Vigreux column, until the boiling point reaches 183°. A few crystals of KMnO<sub>4</sub> are added, and after refluxing, the distillation is continued. [Andrews & Keefer *J Am Chem Soc* **83** 3708 1961, *Beilstein* **2** H 299, **2** I 130, **2** II 263, **2** III 663, **2** IV 868.]

**Valeronitrile** [110-59-8] **M 83.1, b 142.3°/~760mm, d<sub>4</sub><sup>20</sup> 0.799, n<sub>D</sub><sup>15</sup> 1.39913, n<sub>D</sub><sup>30</sup> 1.39037**. Wash the nitrile with half its volume of concentrated HCl (twice), then with saturated aqueous NaHCO<sub>3</sub>, dry it with MgSO<sub>4</sub> and fractionally distil it from P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **2** H 301, **2** I 131, **2** II 267, **2** III 675, **2** IV 875.]

**Vinyl acetate** [108-05-4] **M 86.1, b 72.3°, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>20</sup> 1.396**. Inhibitors such as hydroquinone and other impurities are removed by drying with CaCl<sub>2</sub> and fractionally distilling under nitrogen, then refluxing briefly with a small amount of benzoyl peroxide and redistilling it under nitrogen. Store it in the dark at 0°. Add inhibitor (~0.004%) for storage. [*Beilstein* **2** IV 176.]

**Vinyl butoxyethyl ether (ethylene glycol butyl vinyl ether)** [4223-11-4] **M 144.2, b 70-72°/20mm, d<sub>4</sub><sup>20</sup> 0.866, n<sub>D</sub><sup>20</sup> 1.4220**. Wash the ether with aqueous 1% NaOH, dry with CaH<sub>2</sub>, then reflux with, and distil it from sodium. Stabilise it with 0.5% of 2,6-di-*tert*-butyl-*p*-cresol for storage. [*Beilstein* **1** IV 2387.] **IRRITANT**.

**Vinyl chloroformate** [5130-24-5] **M 106.5, b 46.5°/80mm, 67-69°/atm, 109-110°/760mm, d<sub>4</sub><sup>20</sup> 1.136, n<sub>D</sub><sup>21</sup> 1.420**. It has been fractionated through a Todd column (Model A with ~60 plates) under atmospheric pressure and the purity can be checked by gas chromatography. Stabilise it with 0.5% of 2,6-di-*tert*-butyl-*p*-cresol. It has IR with  $\nu_{\max}$  at 3100 + 2870 (CH<sub>2</sub>), 1780 (C=O), 1640 (C=C) and 940 (CH<sub>2</sub> out-of-plane) and 910 (CH<sub>2</sub> wagging) cm<sup>-1</sup>. [IR: Lee *J Org Chem* **30** 3943 1965, Levallant *Ann Chim (Paris)* **6** 504 1936.] It is used for protecting NH<sub>2</sub> groups in peptide synthesis [Olofson et al. *Tetrahedron Lett* 1563 1977]. [*Beilstein* **3** III 28.]

**Vinyl stearate** [111-63-7] **M 310.5, m 35°, b 166°/1.5mm, 187-188°/4.3mm, d<sub>4</sub><sup>40</sup> 0.8517, n<sub>D</sub><sup>40</sup> 1.4423**. Distil the ester in a vacuum under nitrogen, then crystallise it from acetone (3ml/g) or ethyl acetate at 0°. Store it under nitrogen in the dark. [Swern & Jordan *J Am Chem Soc* **70** 2338 1948, Swern & Jordan *Org Synth* **30** 108 1950, *Beilstein* **2** III 1019.]

## ALICYCLIC COMPOUNDS

**Abietic acid** [514-10-3] **M 302.5, m 172-175°**,  $[\alpha]_{\text{D}}^{25} -116^{\circ}$  ( $-106^{\circ}$ ) (c 1, EtOH), **pK<sup>25</sup> 5.27**. Crystallise it by dissolving 100g of acid in 95% EtOH (700ml), adding to H<sub>2</sub>O (600ml) and cooling. Filter, dry it in a vacuum (over KOH or CaSO<sub>4</sub>) and store it in an O<sub>2</sub>-free atmosphere. It can also be purified *via* the anhydride, tritylabietate and the potassium, piperidine and brucine salts.  $\lambda_{\text{max}}$ : nm(log  $\epsilon$ ): 2343(4.3), 241(4.4), 2505(4.2), 235(4.34) and 240(4.36) in EtOH. [Harris & Sanderson *Org Synth Coll Vol IV 1 1963, J Am Chem Soc 35 3736 1949*, Lambard & Frey *Bull Soc Chim Fr 1194 1948*, Buchbauer et al. *Monatsh Chem 116 1345 1985*.] [*Beilstein 9 IV 2175*.]

**S-Abseic acid (Dormin)** [21293-29-8] **M 264.3, m 160-161°, 161-163° (sublimation)**,  $[\alpha]_{287} +24,000^{\circ}$ ,  $[\alpha]_{245} -69,000^{\circ}$  (c 1-50 $\mu$ g/ml in acidified MeOH or EtOH), **pK<sub>Est</sub> ~3.9**. Crystallise the acid from CCl<sub>4</sub>/petroleum ether, EtOH/hexane and sublime it at 120°. Also purify it by dissolving ~30g in 30ml of EtOAc, adding 100ml of hexane and allow to crystallise overnight (yield 8.4g), **m 156-158°, 161-163°**,  $[\alpha]_{\text{D}}^{20} +426^{\circ}$  (c 0.005M H<sub>2</sub>SO<sub>4</sub> in MeOH). [Cornforth et al. *Nature (London) 206 715 1965*, Soukemp et al. *Helv Chim Acta 72 361 1989*.] The *RS*-isomer was purified on a Kieselgel F<sub>254</sub> plate with toluene/EtOAc/AcOH (50:50:3) and has **m 188-190°** [Cornforth et al. *Aust J Chem 45 179 1992*]. [*Beilstein 17/3 V 13*.]

**Acetylcyclohexane (cyclohexyl methylketone)** [823-76-7] **M 126.2, b 64°/11mm, 76.2-77°/25mm, d<sub>4</sub><sup>20</sup> 0.9178, n<sub>D</sub><sup>20</sup> 1.4519**. Dissolve acetylcyclohexane in Et<sub>2</sub>O, shake it with H<sub>2</sub>O, dry, evaporate and fractionate it under reduced pressure. [UV: Mariella & Raube *J Am Chem Soc 74 518 1952*, enol content: Gero *J Org Chem 19 1960 1954*.] The *semicarbazone* has **m 174°** and the *2,4-dinitrophenylhydrazone* has **m 139-140°** [Theus & Schinz *Helv Chim Acta 39 1290 1956*].

**2-Acetylcyclohexanone** [874-23-7] **M 140.2, m -11°, b 62-64°/2.5mm, 95-98°/10mm, 111-112°/18mm, d<sub>4</sub><sup>20</sup> 1.08, n<sub>D</sub><sup>20</sup> 1.51**. Dissolve it in ligroin (b 30-60°), wash it with saturated aqueous NaHCO<sub>3</sub>, dry over Drierite and fractionate in a vacuum. [Perfetti & Levine *J Am Chem Soc 75 626 1953*, Manyik et al. *J Am Chem Soc 75 5030 1953*, Eistert & Reiss *Chem Ber 87 108 1954*.] It forms a *Cu salt* which crystallises in green leaflets from EtOH, **m 162-163°** [UV: McEntee & Pinder *J Chem Soc 4419 1957*].

**2-Acetylcyclopentanone** [1670-46-8] **M 126.2, b. 72-75°/8mm, 82-86°/12mm, 88°/18mm, d<sub>4</sub><sup>20</sup> 1.043, n<sub>D</sub><sup>20</sup> 1.490**. Dissolve the ketone in petroleum ether (b 30-60°), wash it with saturated aqueous NaHCO<sub>3</sub>, dry over Drierite and fractionate in a vacuum. It gives a violet colour with ethanolic FeCl<sub>3</sub> and is only slowly hydrolysed by 10% aqueous KOH but rapidly on boiling to yield 6-oxoheptanoic acid. [Manyik et al. *J Am Chem Soc 75 5030 1953*, Acheson *J Chem Soc 4232 1956*, UV: Martin & Frenelius *J Am Chem Soc 81 2342 1959*.] It gives a gray green *Cu salt* from Et<sub>2</sub>O/pentane, **m 237-238°** [House & Wasson *J Am Chem Soc 79 1488 1957*].

**2-Acetyl-5,5-dimethylcyclohexane-1,3-dione (2-acetyldimedone)** [1755-15-3] **M 182.2, m 36°, 36-40°, b 132-133°/20mm, 138°/23mm, pK ~4.5**. It can be purified by fractional distillation. *Alternatively*, convert it into the insoluble *Cu salt* in H<sub>2</sub>O and recrystallise it from EtOH, **m ~260°**. The *Cu salt* is decomposed with N H<sub>2</sub>SO<sub>4</sub>, extracted into Et<sub>2</sub>O, dried (Na<sub>2</sub>SO<sub>4</sub>), evaporated and distilled in a vacuum. The residual oil, which solidifies on cooling can be recrystallised from AcOH. It gives a red colour with Fe<sup>3+</sup> ions. The *oxime* has **m 115°**(dec, from EtOH), and with concentrated NH<sub>3</sub> it forms the *mono-imide* which crystallises from H<sub>2</sub>O in needles with **m 133°**. Its UV in EtOH has  $\lambda_{\text{max}}(\epsilon)$  231(10,620) and 273(10,800)nm. It is used as a protecting group for primary amines. [Dieckmann & Stein *Chem Ber 37 3379 1904*, Birch *J Chem Soc 3026 1951*, Crossley & Renouf *J Chem Soc 101 1529 1912*, Nash et al. *Tetrahedron Lett 37 2625 1996*, Kellam et al. *Tetrahedron Lett 38 4849 1997*, *Beilstein 7 H 860, 7 I 471, 7 IV 2756*.]

**4-Acetyl-1-methyl-1-cyclohexene** [6090-09-1] **M 138.2, b 73-75°/7.5mm, 85-86°/13mm, 94-94.7°/20mm, 204.5-206°/747mm, d<sub>4</sub><sup>20</sup> 1.0238, n<sub>D</sub><sup>20</sup> 1.469**. Purify it by fractionation under reduced pressure *in vacuo*, and if it is almost pure it can be fractionated at atmospheric pressure, preferably in an inert atmosphere. It forms two *semicarbazones* one of which is more soluble in \*C<sub>6</sub>H<sub>6</sub>, and both can be recrystallised from EtOH; the more

soluble has **m** 149°(151°), and the less soluble has **m** 172-175°(191°). The 4-nitrophenylhydrazone has **m** 166-167° and the 2,4-dinitrophenylhydrazone has **m** 114-115°. [Pfau & Plattner *Helv Chim Acta* **17** 129, 142 1934, Adler & Vogt *Justus Liebigs Ann Chem* **564** 109 1949.]

**2-Acetyl-1-methyl-3,5-dioxo-1-methylcyclohexanecarboxylic acid (ADCC-linker)** [181486-37-3] **M 212.2, m 95-99°, pK<sub>Est</sub> ~4.5**. It is prepared from 3,5-dioxo-1-methylcyclohexane-1-carboxylic acid methyl ester (obtained from 3,5-dimethoxybenzoic acid [119-52-8] *via* Birch reduction, methylation with MeI and treatment with aqueous HCl) by acetylation to 3-acetoxy-5-oxo-1-methylcyclohex-3,4-ene-1-carboxylic acid methyl ester, rearrangement (heating with DMAP) to 4-acetyl-3,5-dioxo-1-methylcyclohexane carboxylic acid methyl ester (93% yield) followed by hydrolysis (LiOH/THF/MeOH/H<sub>2</sub>O) and H<sup>+</sup> ion-exchange purification to give the ADCC-linker as a crystalline solid. It should be stored at ~5°. After attaching to a solid support through the carboxy function (e.g. to amino-modified polystyrene beads), it is used for linking to primary amines (including  $\alpha$ -amino-acid esters) *via* enamine formation of the 4-acetyl group for combinatorial synthesis. The linker is stable to acids such as CF<sub>3</sub>CO<sub>2</sub>H, bases such as piperidine or BTU, and uronium type coupling agents; and by treatment with 2% hydrazine in DMF, the primary amine is released quantitatively from the support. Its <sup>1</sup>H NMR (250MHz, Me<sub>2</sub>SO) has  $\delta$  at 1.25 (s, Me-C(1)), 2.50 (s, MeCO-C(4)), 2.50-3.00 (m, 2H-C(3), 2H-C(5)), 17.80 (2H, enol OH and COOH). [Bannwarth et al. *Bioorg Med Chem Lett* **6** 1525 1966.]

**Adamantane (tricyclo[3.3.1.1<sup>3,7</sup>]-decane)** [281-23-2] **M 136.2, m 269.6-270.8° (sublimes)**. Crystallise adamantane from acetone or cyclohexane, and sublime it in a vacuum below its melting point [Butler et al. *J Chem Soc, Faraday Trans I* **82** 535 1986]. Adamantane is also purified by dissolving it in *n*-heptane (*ca* 10ml/g of adamantane) on a hot plate, adding activated charcoal (2g/100g of adamantane), and boiling for 30 minutes, filtering the hot solution through a filter paper, concentrating the filtrate until crystallisation just starts, adding one quarter of the original volume of *n*-heptane, and allowing to cool slowly over a period of hours. The supernatant is decanted off and the crystals are dried *in vacuo* at 25°. [Prelog & Seiwerth *Chem Ber* **74** 1769 1941, Schleyer et al. *Org Synth Coll Vol V* 16 1973, Walter et al. *J Am Chem Soc* **107** 793 1985.] [*Beilstein* **5** III 393, **5** IV 469.]

**1-Adamantane acetic acid** [4942-47-6] **M 194.3, m 136°, pK<sub>Est</sub> ~4.8**. Dissolve the acid in hot N NaOH, treat with charcoal, filter and acidify. Collect the solid, wash it with H<sub>2</sub>O, dry and recrystallise it from MeOH. [Stetter et al. *Chem Ber* **92** 1629 1959.] The *acid chloride* [2094-72-6] has **M** 168.7, **m** 51-54°, and **b** 135-136°/1mm. **LACHRYMATORY**.

**1-Adamantane carboxylic acid** [828-51-3] **M 180.3, m 175-176.5°, 177°, pK<sub>Est</sub> ~4.9**. Possible impurities are trimethylacetic acid and C9 and C13 acids. Dissolve 15g of the acid in CCl<sub>4</sub> (300ml) and shake with 110ml of 15N aqueous NH<sub>3</sub> whereby the ammonium salt separates and is collected. Acid impurities form soluble ammonium salts. The salt is washed with cold Me<sub>2</sub>CO (20ml) and suspended in H<sub>2</sub>O (250mL). This is treated with 12N HCl and extracted with CHCl<sub>3</sub> (100ml). The dried (Na<sub>2</sub>SO<sub>4</sub>) extract is evaporated and the residue is recrystallised from a mixture of MeOH (30ml) and H<sub>2</sub>O (*ca* 10ml) to give the pure acid (10-11g). [Koch & Haaf *Org Synth Coll Vol V* 20 1973.] It was also recrystallised from absolute EtOH and dried under vacuum at 100°.

*Alternatively*, the acid (5g) is refluxed for 2 hours with 15ml of MeOH and 2ml of 98% H<sub>2</sub>SO<sub>4</sub> (cool when mixing this solution). Pour into 10 volumes of H<sub>2</sub>O and extract with the minimum volume of CHCl<sub>3</sub> to give clear separation of phases. The extract is washed with H<sub>2</sub>O, dried (CaCl<sub>2</sub>) and distilled. The *methyl ester* is collected at 77-79°/1mm, **m** 38-39°. The ester is hydrolysed with the calculated amount of N KOH and refluxed until clear. Acidification with HCl provides the pure acid with 90% recovery. The *amide* crystallises from cyclohexane, **m** 189°. [Stetter et al. *Chem Ber* **92** 1629 1959.] [*Beilstein* **9** IV 253.]

**1,3-Adamantane diamine dihydrochloride** [26562-81-2] **M 239.2, m >310°, pK<sub>Est(1)</sub> ~8.1, pK<sub>Est(2)</sub> ~10.1**. Dissolve it in boiling conc HCl (400mg in 15ml) and evaporate to dryness. Dissolve it in absolute EtOH and add dry Et<sub>2</sub>O to crystallise the *dihydrochloride*. [Stetter & Wulff *Chem Ber* **93** 1366 1960, *Beilstein* **13** III 27,]

**1,3-Adamantane dicarboxylic acid** [39269-10-8] **M 224.3, m 276°, 276-278°, 279°, pK<sub>Est(1)</sub> ~4.9, pK<sub>Est(2)</sub> 5.9**. Dissolve the acid in aqueous NaOH, treat with charcoal, filter and acidify with dilute HCl. It crystallises

from MeOH. [Stetter & Wulff *Chem Ber* **93** 1366 1960, *Beilstein* **9** III 4066, **9** IV 2997.]

**1-Adamantane methylamine** [17768-41-1] **M 165.3, b 83-85°/0.3mm, d<sub>4</sub><sup>20</sup> 0.93, pK<sub>Est</sub> ~10.2.** Dissolve the amine in Et<sub>2</sub>O, dry over KOH and distil it. The *N*-Tosyl derivative has **m 134-135°** (from EtOH). [Stetter & Goebel *Chem Ber* **96** 550 1963.]

**1-Adamantanol (1-hydroxyadamantane)** [768-95-6] **M 152.4, m 288.5-290°.** If 2-adamantanol is a suspected impurity, then dissolve the substance (10g) in acetone (100ml) and add Jones's reagent [CrO<sub>3</sub> (10.3g) in H<sub>2</sub>O (30ml)], then conc H<sub>2</sub>SO<sub>4</sub> (8.7ml) is added dropwise (turns green in colour) until excess reagent is present (slight red colour). Stir overnight, decant the acetone solution from the Cr salts and adamantan-2-one, dry (Na<sub>2</sub>SO<sub>4</sub>) and evaporate to dryness. The residue (*ca* 7g) is chromatographed through Al<sub>2</sub>O<sub>3</sub> (250g) and washed with 50% \*benzene/petroleum ether (b 40-60°), then 100% Et<sub>2</sub>O (to remove any adamantan-2-one present) and the 1-adamantanol is then eluted with 5% MeOH in Et<sub>2</sub>O. The eluate is evaporated, and the residue is recrystallised from petroleum ether (b 30-60°) at -70°, **m 287.2-288.5°.** It also crystallises from MeOH and can be sublimed *in vacuo*. It has characteristic IR, with  $\nu_{\max}$  at 3640, 1114, 1086, 982 and 930 cm<sup>-1</sup>. [Schleyer & Nicholas *J Am Chem Soc* **83** 182 1961.] [*Beilstein* **6** IV 391.]

*Alternatively*, if free from the 2-isomer, dissolve it in tetrahydrofuran, and dilute with H<sub>2</sub>O to precipitate the alcohol. Collect, dry and sublime it in a vacuum at 130°. [Stetter et al. *Chem Ber* **92** 1629 1959.]

**2-Adamantanol (2-hydroxyadamantane)** [700-57-2] **M 152.4, m 296.2-297.7°.** It can be purified by chromatography as for the 1-isomer. It crystallises from cyclohexane and has characteristic IR with  $\nu_{\max}$  at 3600, 1053, 1029 and 992 cm<sup>-1</sup> [Schleyer & Nicholas *J Am Chem Soc* **83** 182 1961.]

**2-Adamantanone** [700-58-3] **M 150.2, m 256-258°(sublimes).** Purify 2-admantanone by repeated sublimation *in vacuo*. [Butler et al. *J Chem Soc, Faraday Trans II* **82** 535 1986.]

***N*-(1-Adamantyl)acetamide** [880-52-4] **M 193.3, m 149°.** Wash the amide well with H<sub>2</sub>O, dry and recrystallise it from cyclohexane. *It is an irritant.* [Stetter et al. *Chem Ber* **92** 1629 1959.]

**1-Adamantylamine (1-adamantanamine)** [768-94-5] **M 151.2, m 160-190° (sealed tube), 180-192°, 208-210°, pK<sub>25</sub> 10.58.** Dissolve the amine in Et<sub>2</sub>O, dry it over KOH, evaporate and sublime it *in vacuo*. [Stetter et al. *Chem Ber* **93** 226 1960.]

**1-Adamantylamine hydrochloride (Amantadine hydrochloride, Amazolone, Mantadan, Mantadix, Virofral)** [665-66-7] **M 187.7, m >360° (dec), pK<sub>25</sub> 10.58.** Dissolve the salt in dry EtOH, add a few drops of dry EtOH saturated with HCl gas, followed by dry Et<sub>2</sub>O to crystallise the hydrochloride out. Dry the salt in a vacuum. Its solubility in H<sub>2</sub>O is >5%; and it is soluble in EtOH but insoluble Et<sub>2</sub>O. [Stetter et al. *Chem Ber* **93** 226 1960.] It is antiviral and anti-Parkinsonian agent [Kornhuber et al. *J Neural Transm* **46** (Suppl) 399 1995].

**2-Adamantylamine hydrochloride** [10523-68-9] **M 187.7, m >300°, pK<sub>Est</sub> ~10.4.** The *free amine* in Et<sub>2</sub>O, liberated by the action of alkali in H<sub>2</sub>O, is dried over KOH, filtered, evaporated and sublimed at 110°/12torr, **m 230-236°.** The base is dissolved in EtOH, sufficient ethanolic HCl is added dropwise and crystallised by the addition of Et<sub>2</sub>O. Dry it *in vacuo*. [Stetter et al. *Justus Liebigs Ann Chem* **658** 151 1962.]

**1-Adamantyl bromide** [768-90-1] **M 215.1, m 117-119°, 118°, 119.5-120°.** If coloured, dissolve it in CCl<sub>4</sub>, wash with H<sub>2</sub>O, treat with charcoal, dry (CaCl<sub>2</sub>), filter and evaporate to dryness. Dissolve the residue in a small volume of MeOH and cool in a CO<sub>2</sub>/trichloroethylene bath and collect the crystals. Sublime it at 90-100°/water pump vacuum. [Stetter et al. *Chem Ber* **92** 1629 1959, Schleyer & Nicholas *J Am Chem Soc* **83** 2700 1961, *Beilstein* **5** III 469.]

**1-Adamantyl bromomethylketone** [5122-82-7] **M 257.2, m 76-79°, 78-79°.** Dissolve the ketone in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dry (MgSO<sub>4</sub>), evaporate and crystallise the residue from small volumes of MeOH. **LACHRYMATORY.** [Stetter & Rauscher *Chem Ber* **93** 2054 1960.]

**1-Adamantyl chloride** [935-56-8] **M 170.7, m 164.3-165.6°**. Crystallise the chloride from aqueous MeOH and sublime it at 100°/12torr. It also crystallises from MeOH at 70°. Do not keep in contact with MeOH for too long. [Stetter et al. *Chem Ber* **92** 1629 1959, Schleyer & Nicholas *J Am Chem Soc* **83** 2700 1961, *Beilstein* 5 IV 469.]

**1-Adamantyl chloroformate** [5854-52-4] **M 214.6, m 46-47°**. Crystallise it from petroleum ether (b 30-60°) at -20°. Also purify it as for 1-adamantyl fluoroformate below. Its IR has  $\nu_{\max}$  at 4.2, 5.6 and 8.4 $\mu$  (2380, 1786 and 1190  $\text{cm}^{-1}$ ). [Haas et al. *J Am Chem Soc* **88** 1988 1966, cf Moroder et al. *Hoppe-Seyler's Z Physiol Chem* **357** 1647 1976.]

**1-(1-Adamantyl)ethylamine hydrochloride (Rimantadine hydrochloride, Flumadine, Roflual)** [1501-84-4] **M 215.8, m >300° , 373-375°(sealed tube), pK<sub>Est</sub> ~10.4**. It is prepared by adding the oxime of 1-adamantyl methyl ketone (6 parts, see [1660-04-4] below) in dry THF (200 parts) to a cold suspension of excess of LiAlH<sub>4</sub> in Et<sub>2</sub>O, refluxing for 1 hour, decomposing with cold brine, making strongly alkaline and extracting thoroughly with Et<sub>2</sub>O. The dried (Na<sub>2</sub>CO<sub>3</sub>) extract, is filtered, concentrated, and saturated with dry HCl. The hydrochloride (~5.25 parts) is collected, dried and recrystallised as for 2-adamantylamine hydrochloride above. [US patent to du Pont de Nemours 1069563 (1969, Brit amended), *Chem Abstr* **75** 140372w 1971.] It is an antiviral agent [Burkinskaya et al. *J Gen Virol* **60** 49 1982].

**1-Adamantyl fluoride (1-fluoroadamantane)** [768-92-3] **M 154.2, m 210-212°(dec, sealed tube), 259-260°(dec)**. Dissolve it in Et<sub>2</sub>O, dry over Na<sub>2</sub>SO<sub>4</sub>, evaporate to dryness and sublime the residue at 90-100°/12mm. Recrystallise the sublimate from MeOH, **m 259-260°**. To remove 1-hydroxyadamantane impurity, dissolve it in cyclohexane, cool for many hours, filter off the hydroxyadamantane, and evaporate to dryness, or by passage through an Al<sub>2</sub>O<sub>5</sub> column in dry cyclohexane. Recrystallise the residue from petroleum ether at -77° and sublime it in vacuum, **m 210-212°dec (sealed tube)**. [Bhandari & Pinock *Synthesis* 655 1974, NMR: Fort et al. *J Org Chem* **30** 789 1965.]

**1-Adamantyl fluoroformate** [62087-82-5] **M 198.2, m 31-32°**. Dissolve it in *n*-hexane (ca 10g in 150 ml) and keep at 0° for 24 hours. Any 1-adamantanol present will separate. Filter and evaporate to dryness. The crystalline residue has **m 31-32°** and is recrystallised from *n*-hexane (90g/500ml), (IR (KBr):  $\nu_{\max}$  1242, 1824 and 2340  $\text{cm}^{-1}$ ). There should be no OH str band above 2500  $\text{cm}^{-1}$ . [Moroder et al. *Hoppe-Seyler's Z Physiol Chem* **357** 1647 1976, cf Haas et al. *J Am Chem Soc* **88** 1988 1966.]

**1-Adamantyl iodide (1-iodoadamantane)** [768-93-4] **M 262.1, m 75.3-76.4°**. Dissolve the iodide in Et<sub>2</sub>O, shake with aqueous NaHSO<sub>3</sub>, aqueous K<sub>2</sub>CO<sub>3</sub>, and H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and recrystallise it from MeOH at -70° (to avoid alcoholysis) to give white crystals. [Schleyer & Nicholas *J Am Chem Soc* **83** 2700 1961, lit **m** of 151-152.5° is incorrect.] Also purify by recrystallisation from petroleum ether (40-60°C) followed by rigorous drying and repeated sublimation. [*Beilstein* 5 IV 470.]

**1-Adamantyl isocyanate** [4411-25-0] **M 177.3, m 144-145°**. Recrystallise the isocyanate from *n*-hexane and sublime it. **Irritant**. [Stetter & Wulff *Chem Ber* **95** 2302 1962.]

**1-Adamantyl isothiocyanate** [4411-26-1] **M 193.3, m 168-169°**. Dissolve it in Et<sub>2</sub>O, wash with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and sublime the residue in a vacuum at 140°, then recrystallise it from MeOH. **Irritant**. [Stetter & Wulff *Chem Ber* **95** 2302 1962.]

**1-Adamantyl methyl ketone** [1660-04-4] **M 178.3, m 53-55°, 54-55°**. The ketone is prepared by bubbling acetylene through a vigorously stirred solution of 1-adamantylbromide (1g, see [769-90-1]) in 96% H<sub>2</sub>SO<sub>4</sub> (40ml) at 5° for 5 hours with evolution of HBr. The mixture is poured onto ice, extracted with Et<sub>2</sub>O, the extract is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated to give an ~80% yield of ketone. This was purified *via* the 2,4-dinitrophenyl hydrazine derivative (by passage of a solution through silica gel/\*C<sub>6</sub>H<sub>6</sub>) which gave orange needles from AcOH/EtOH or EtOH with **m 221° (219-222° and 220-222° also reported)**. The ketone has an IR (KBr) peak at  $\nu_{\max}$  1690  $\text{cm}^{-1}$  (C=O); its <sup>1</sup>HNMR has  $\tau$  at 7.96 (s, 3H), 7.90 and 8.30 (15H); and its MS has *m/z* at 178 (M<sup>+</sup>). [Kell & McQuillin *J Chem Soc Perk I* 2100 1972, Sasaki et al. *Chem Commun* 780 1968.]

Alternatively, 1-(ethoxycarbonylmethylcarbonyl)adamantane [1-(EtOCOCH<sub>2</sub>CO)-Ad], b 108-115°/0.06mm, prepared from 1-adamantylcarbonyl chloride and diethyl malonate] (~25g) was hydrolysed and decarboxylated by boiling in AcOH (50ml), H<sub>2</sub>O (30ml) and concentrated H<sub>2</sub>SO<sub>4</sub> (5.5ml) until evolution of CO<sub>2</sub> ceased, poured into ice-water (~300ml), the oily *ketone* solidified (94-96% yield), and was recrystallised from MeOH or aqueous MeOH. It sublimes at 40°/0.1mm. **1-Adamantyl methyl ketone oxime**, m 182-184° (used for the preparation of *Rimantadine hydrochloride* see [1501-84-4] above) is obtained by shaking hydroxylamine hydrochloride (2.5g) and recrystallised NaOAc (4g) in H<sub>2</sub>O (10ml) in a glass test tube, then warming to 40° and adding the *ketone* (2.5g), replacing the stopper and shaking vigorously for a few minutes. The crystalline oxime that soon separates is filtered off, washed with H<sub>2</sub>O, and gives colourless plates in high yield upon recrystallising from aqueous dioxane. [Stetter & Rauscher *Chem Ber* **93** 2054 1960, see also Hála & Landa *Col Czech Chem Comm* **25** 2692 1960.]

**N-(1-Adamantyl)urea** [13072-69-0] M 194.2, m >250°(dec), 268-272°(dec). Wash the urea with H<sub>2</sub>O and dioxane and recrystallise it from EtOH. [Stetter & Wulff *Chem Ber* **95** 2302 1962.]

(-)-**Alloaromadendrene** [25246-27-9] M 204.4, b 96°/2mm, 265-267°/atm, [α]<sub>D</sub><sup>25</sup>-22° (neat), d<sub>4</sub><sup>20</sup> 0.923, n<sub>D</sub><sup>23</sup> 1.501. Fractionally distil it from Na. It has IR bands at 6.06 and 11.27μ due to C=CH<sub>2</sub>. [Birch *J Chem Soc* 715 1953, cf Büchi et al. *J Am Chem Soc* **91** 6473 1969.]

**Cis-(±)-(1-*RS*,2-*SR*)-6-Amino-3-cyclohexene-1-carboxylic acid (cis-(±)-1,2,3,6-tetrahydroanthranilic acid)** [54162-90-2] M 141.5, m 216-218°, pK<sub>Est(1)</sub> ~3.5, pK<sub>Est(2)</sub> ~10.2. Purify the free amino-acid by dissolving it in H<sub>2</sub>O and passing it through a Dowex 50W (acid form) column and eluting with 1M aqueous NH<sub>4</sub>OH. The eluate is evaporated (*in vacuo*) and the residue is dissolved in H<sub>2</sub>O. Me<sub>2</sub>CO is added to turbidity, cooled at 0° and the colourless crystals of the amino-acid are collected and dried *in vacuo* [Bernáth et al. *Tetrahedron* **41** 1315 1958, cf Mazza & Crapetta *Gazzetta* **57** 297 1927]. In earlier work, it was recrystallised from aqueous EtOH and had reported melting points of 265-265° [Kricheldorf *Justus Liebigs Ann Chem* 1378 1975] and 269-271° [Marconi & Mazzochi *J Org Chem* **31** 1372 1966]. The *hydrochloride* [57266-56-5] M 177.6 has m 210-213° and the *methyl ester hydrochloride* [52766-61-2] has m 85-87° (from Et<sub>2</sub>O). The *trans-(±)-(1RS,2RS)-amino-acid* [97945-19-2] crystallises from aqueous Me<sub>2</sub>CO with m 267-269°. [Beilstein **14** II 203.]

**α-Amyrin** [638-95-9] M 426.7, m 186°, 244°/0.8mm, [α]<sub>D</sub><sup>25</sup> +85° (c 2, CHCl<sub>3</sub>). Purify it by acetylation to the acetate followed by hydrolysis and recrystallisation from aqueous MeOH or from EtOH. The *acetate* when crystallised from petroleum ether, *n*-heptane or CHCl<sub>3</sub>/MeOH, and sublimed *in vacuo* has m 227° (225-226°) and [α]<sub>D</sub><sup>20</sup> +76.4° (c 0.6, CHCl<sub>3</sub>). [Bently et al. *J Chem Soc* 3672 1953, IR: Cole & Thornton *J Chem Soc* 1332 1957, Corey & Cantrall *J Am Chem Soc* **81** 1745 1958, Beilstein **6** III 2889, **6** IV 4191.]

**β-Amyrin** [508-04-3] M 426.7, m 197-197.5°, 204-205°, 260°/0.8mm, [α]<sub>D</sub><sup>23</sup> +91° (c 0.9, CHCl<sub>3</sub>). Purify it through an Al<sub>2</sub>O<sub>3</sub> column and elute with petroleum ether (40-60°) then Et<sub>2</sub>O and recrystallise from petroleum ether or EtOH. The *acetate* crystallises from Ac<sub>2</sub>O or petroleum ether and has m 242-143° and [α]<sub>D</sub><sup>20</sup> +82.8° (c 0.81, CHCl<sub>3</sub>) [Crow & Michael *Aust J Chem* **8** 133 1955, Barton et al. *J Chem Soc (C)* 1031 1968, Beilstein **6** III 1894, **6** IV 4195.]

**1,1'-Azobis(cyclohexane carbonitrile)** [2094-98-6] M 244.3, m 114-114.5°, 114-115°, ε<sub>350nm</sub> 16.0. Purify the nitrile by dissolving it in boiling 95%EOH as rapidly as possible, cool overnight at 0°, filter, wash with a little EtOH and dry it in a vacuum desiccator over CaCl<sub>2</sub>. Note that prolonged heating >80° causes decomposition. Recrystallise it from EtOH. It should be regarded as **potentially explosive**. It is a radical initiator. [Overberger et al. *Org Synth Coll Vol* **IV** 66 1963, *J Am Chem Soc* **71** 2661 1949, Beilstein **16** II 97.]

**Bicyclohexyl** [92-51-3] M 166.3, b 238° (*cis-cis*), 217-219° (*trans-trans*). Shake bicyclohexyl repeatedly with aqueous KMnO<sub>4</sub> and with conc H<sub>2</sub>SO<sub>4</sub>, wash it with water, dry, first with CaCl<sub>2</sub> then with sodium, and distil it. [Mackenzie *J Am Chem Soc* **77** 2214 1955, Beilstein **5** IV 334.]

**Bicyclo[3.2.1]octane** [6221-55-2] M 110.2, m 141°. Purify it by zone melting. It has been sublimed under N<sub>2</sub>

at 70° and atmospheric pressure (closed vessel), and resublimed over P<sub>2</sub>O<sub>5</sub> to give an analytically pure sample **m** 137.5-139.5°. [Von E Doering & Farber *J Am Chem Soc* **71** 1514 1949, Cope et al. *J Am Chem Soc* **82** 4299 1960, NMR: Stothers et al. *Can J Chem* **55** 841 1977.]

**1R-2-endo-Borneol** [464-43-7] **M 154.3, m 204.5-205.5°, 208°, 212°/atm, [α]<sub>D</sub><sup>20</sup> +37° (c 5, EtOH)**. It can be steam distilled, the distillate is extracted into Et<sub>2</sub>O, the extract dried with Drierite, filtered and evaporated. The residue is then recrystallised from boiling EtOH (charcoal) or petroleum ether. [Clark & Read *J Chem Soc* 1773 1934, *Beilstein* **6** III 295, **6** IV 281.]

**(±)-Borneol** [6627-72-1] **M 154.3, m 206-207°, 210-215°**. Crystallise borneol from petroleum ether (b 60-80°) and sublime it *in vacuo*. [*Beilstein* **6** II 81, **6** IV 281.]

**3-Bromoadamantane-1-carboxylic acid** [21816-08-0] **M 259.1, m 145-146°, 146.5°, 147-150°, pK<sup>25</sup> 6.28 (50% aqueous EtOH)**. Purify the acid by recrystallising it from cyclohexane and/or subliming at 130°/10mm. It can be converted to the *methyl ester* (diazomethane) with **m** 32° (from petroleum ether at -10°). [Stetter & Mayer *Chem Ber* **95** 667 1962, Stetter & Wulff *Chem Ber* **93** 1366 1960, Bayal & Lantvoev *J Org Chem USSR (Engl Trans)* **9** 291 1973.]

**(+)-3-Bromocamphor-8-sulfonic acid** [5344-58-1] **M 311.2, m 195-196°(anhydrous), [α]<sub>D</sub><sup>20</sup> +88.3° (c 1, H<sub>2</sub>O), pK ~0**. Crystallise the acid from water. The *ammonium salt* has **m** 268-207°, [α]<sub>D</sub><sup>20</sup> +81.9° (c 2.2, H<sub>2</sub>O). [Kauffman *J Prakt Chem* **33** 95 1966.]

**1R(endo, anti)-3-Bromocamphor-8-sulfonic acid ammonium salt**, see entry [55870-50-3] on p. 504.

**(+)-3-Bromocamphor-10-sulfonic acid hydrate** [67999-30-8] **M 329.2, m 119-121°, [α]<sub>D</sub><sup>20</sup> +98.3° (c 1, H<sub>2</sub>O), pK ~0**. Crystallise the acid from water. [Boyle *Quart Rev Chem Soc* **25** 323 1971, UV: Lowry & Owen *J Chem Soc* 609 1926, *Beilstein* **11** II 181, **11** III 592.]

**4-tert-Butyl-1-cyclohexanone** [98-53-3] **M 154.3, m 49-50°, 52-52.5, b 90-92°/9mm**. Purify it *via* the *semicarbazone* (crystallised from EtOH with **m** 203-205°), hydrolyse this with dilute HCl and steam distil it. The distillate is extracted into Et<sub>2</sub>O, dried, evaporated and the residue is recrystallised from pentane, aqueous EtOH or EtOH [Houlihan *J Org Chem* **27** 3860 1962]. The *oxime* recrystallises from 1,2-dichloropropane and has **m** 137.5-138.5°. [Harvill et al. *J Org Chem* **15** 58 1950, *Beilstein* **7** IV 82.]

**(+)-Calarene** [(+)-β-gurjunen, 1,3,3,11-tetramethyltricyclo[5.4.0.0<sup>2,4</sup>]undecan-7-ene, (1aR)-1,1,7c,7ac-tetramethyl-1a,2,3,5,6,7,7a,7b-octahydro-1H-cyclopropa[α]naphthalene, new name 1(10)aristolene] [17334-55-3] **M 204.35, b 45-47°/0.008-0.01mm, 255-258°/atm, d<sub>4</sub><sup>20</sup> 0.9340, n<sub>D</sub><sup>20</sup> 1.55051, [α]<sub>D</sub><sup>20</sup> +73° (c 2, EtOH), +81.8° (neat)**. Purify the sesquiterpene Calarene by gas chromatography (7% propylene glycol adipate on unglazed tile particles of size 0.2-0.3mm, 400 cm column length and 0.6 cm diameter, at 184°, with N<sub>2</sub> carrier gas at a flow rate of 0.54 ml/sec using a thermal detector). Also purify it by chromatography on alumina (200 times the weight of calarene) and elute with petroleum ether. Its UV has λ<sub>max</sub> at 200 and 210 nm (ε 9560, 5480) in EtOH. [IR: Sorm *Col Czech Chem Comm* **18** 512 1953, **29** 795 1964, Buchi et al. *Tetrahedron Lett* **3** 827 1962, Vrkoc *Tetrahedron Lett* **4** 225 1963, *Beilstein* **5** III 1093.]

**1R,4S(-)-Camphanic acid** [13429-83-9] **M 198.2, m 190-192°, 198-200°, [α]<sub>548</sub><sup>20</sup> -22.5° (c 1, dioxane), -4.4° (c 8, EtOH), pK<sub>Est</sub> ~3.8**. Dissolve it in CH<sub>2</sub>Cl<sub>2</sub>, dry (MgSO<sub>4</sub>), filter, evaporate and the residue is sublimed at 120°/0.5mm or 140°/1mm. [Gerlach *Helv Chim Acta* **61** 2773 1978, *Beilstein* **18/8** V 101.]

**1R,4S(-)-Camphanic acid chloride** [39637-74-6] **M 216.7, m 65-66.5°, 70.5-71°, [α]<sub>548</sub> -23° (c 2, CCl<sub>4</sub>), [α]<sub>364</sub><sup>20</sup> -29.2°, [α]<sub>405</sub><sup>20</sup> -18.0°, [α]<sub>436</sub><sup>20</sup> -13.5°, [α]<sub>546</sub><sup>20</sup> -7.8°, [α]<sub>578</sub><sup>20</sup> -6.0°, (c 0.67, \*C<sub>6</sub>H<sub>6</sub>)**. It is soluble in toluene (50g/100ml at 0°) and crystallises from petroleum ether (b 40-60°). It sublimes at 70°/5mm. Store it dry at 0°, IR (CCl<sub>4</sub>) has ν<sub>max</sub> at 1805s and 1780m cm<sup>-1</sup>. [Armarego et al. *J Chem Soc, Perkin Trans I* 2229 1976, Gerlach *Helv Chim Acta* **51** 1587 1968, Gerlach *Helv Chim Acta* **68** 1815 1985, *Beilstein* **18/8** V 101.]

**RS-Camphene** [565-00-4] **M 136.2, m 51-52°**, **b 40-70°/10mm**. Crystallised it twice from EtOH, then repeatedly melted and frozen at 30mm pressure. [Williams & Smyth *J Am Chem Soc* **84** 1808 1962.] Alternatively, it is dissolved in Et<sub>2</sub>O, dried over CaCl<sub>2</sub> and Na, filtered, evaporated and the residue is sublimed in a vacuum [NMR: Hana & Koch *Chem Ber* **111** 2527 1978].

**(-)-Camphene (1S-2,2-dimethyl-3-methylene norbornane)** [5794-04-7] **M 136.2, m 49.2-49.6°, 49-50°, b 79-80°/58mm, 91.5°/100mm, d<sub>4</sub><sup>54</sup> 0.8412, n<sub>D</sub><sup>54</sup> 1.4564, [α]<sub>D</sub><sup>21</sup> -119.1° (c 2.3, \*C<sub>6</sub>H<sub>6</sub>), -117.5° (c 19, toluene), -113.5° (c 9.7, Et<sub>2</sub>O)**. Purify the norbornane by fractionation through a Stedman column (see p. 11) at 100mm in a N<sub>2</sub> atmosphere, crystallise it from EtOH and sublime it in a vacuum below its melting point. It is characterised by its *camphenilone semicarbazone*, **m 217-218.5°**, or *camphor semicarbazone*, **m 236-238°**. [NMR: Hana & Koch *Chem Ber* **111** 2527 1978, Bartlett et al. *Justus Liebigs Ann Chem* **623** 217 1959, Bain et al. *J Am Chem Soc* **72** 3124 1950, Beilstein **5** IV 461.]

**Camphor (1R-bornan-2-one)** [R-(+)- 464-49-3, S-(-)- 464-48-2] **M 136.2, m 178.8°, 179.97° (open capillary), b 204°/atm, [α]<sub>D</sub><sup>35</sup> (+) and (-) 59.6° (in EtOH), [α]<sub>D</sub><sup>20</sup> (+) and (-) 44.3° (c 10, EtOH), [α]<sub>D</sub><sup>179</sup> (+) and (-) 70.85°(melt)**. Crystallise it from EtOH, 50% EtOH/water, MeOH, or petroleum ether or from glacial acetic acid by addition of water. It can be sublimed (50°/14mm) and also fractionally crystallised from its own melt. It is steam volatile. It should be stored in tight containers as it is appreciably volatile at room temperature. The solubility is 0.1% (H<sub>2</sub>O), 100% (EtOH), 173% (Et<sub>2</sub>O) and 300% (CHCl<sub>3</sub>). The *R-oxime* (from Et<sub>2</sub>O, CHCl<sub>3</sub>, or aqueous EtOH) has **m 119° [α]<sub>D</sub><sup>20</sup> -42.4° (c 3, EtOH)**, the *± oxime* has **m 118-119°**. It has a characteristic odour. [Asahina & Ishidate *Chem Ber* **67** 1432 1934, Allan & Rodgers *J Chem Soc (B)* 632 1971, UV, NMR: Fairley et al. *J Chem Soc, Perkin Trans 1* 2109 1973, White & Bishop *J Am Chem Soc* **62** 8 1940, Beilstein **7** IV 213.]

**Camphoric acid (1,2,2-trimethylcyclopentan-1r,3c-dicarboxylic acid)** [1R,2S)-(+)- 124-83-4, 1S,2R)-(-)- 560-09-8] **M 200.2, m 186-188°, 187°, 186.5-189°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 57° (c 1, EtOH), [α]<sub>D</sub><sup>20</sup> (+) and (-) 47.7° (c 4, EtOH), pK<sub>1</sub><sup>25</sup> 4.71, pK<sub>2</sub><sup>25</sup> 5.83 (for + isomer)**. Purify the acid by re-precipitation from an alkaline solution by HCl, filter it off, and recrystallise it from water several times, rejecting the first crop. It forms leaflets from EtOH, Me<sub>2</sub>CO and H<sub>2</sub>O, and is insoluble in CHCl<sub>3</sub>. Its solubility in H<sub>2</sub>O is 0.8% at 25° and 10% at 100°, 50% in EtOH and 5% in ethylene glycol. The *(±)-acid* has **m 202-203°**. The *(+)-1-methyl ester* has **m 86° (from petroleum ether) [α]<sub>D</sub><sup>20</sup> +45° (c 4, EtOH)**, and the *(+)-3-methyl ester* has **m 77° (from petroleum ether) [α]<sub>D</sub><sup>17.5</sup> +53.9° (c 3, EtOH)**. [Rupe & Thommen *Helv Chim Acta* **30** 933 1947, Tiovonan et al. *Acta Chem Scand* **2** 597 1948, Howell & Fisher *J Am Chem Soc* **80** 6316 1958, Beilstein **9** IV 2851.]

**(±)-Camphoric anhydride** [595-30-2, 76-32-4] **M 182.2, transition temp. 135°, m 223.5°**. Crystallise the anhydride from EtOH. If it contains too much of the acid (check the IR), then reflux it in Ac<sub>2</sub>O, concentrate and collect the crystals, wash them with petroleum ether and dry them *in vacuo*. [Bunton et al *J Chem Soc* 2918 1963, NMR: Baker & Davis *Tetrahedron* **24** 1663 1968, Beilstein **18** H 400, 401.]

**Camphorquinone (borna-2,3-dione)** [1R-(-)- 10334-26-6, 1S-(+)- 2767-84-2] **M 166.2, m 198.7°, 198-199°, 197-201°, [α]<sub>D</sub><sup>25</sup> (-) and (+) 101.1° (c 2, EtOH)**. It can be purified by steam distillation, recrystallisation (yellow prisms) from EtOH, \*C<sub>6</sub>H<sub>6</sub> or Et<sub>2</sub>O/petroleum ether and it can be sublimed in a vacuum. The *(±)-quinone* forms needles from EtOH, **m 197-198°, 203°**. [Buxtorf & Flatt *Helv Chim Acta* **13** 1026 1930, Asahena et al. *Chem Ber* **67** 1432 1934, Beilstein **7** I 325.]

**RS-Camphorquinone** [10373-78-1] **M 166.2, m 199-202°**. Purification is the same as for above enantiomers. [Huckel & Fichtig *Justus Liebigs Ann Chem* **628** 81 1962, Evans et al. *J Chem Soc* 137 1939, Beilstein **7** I 325.]

**(1R)-(-)Camphor-10-sulfonic acid** [35963-20-3] **M 232.3, m 197.4-198°(dec), 197-198°, [α]<sub>D</sub><sup>20</sup> -20.7° (c 5.4, H<sub>2</sub>O), pK<sub>Est</sub> ~ -1**. It forms prisms from AcOH or EtOAc, and is deliquescent in moist air. Store it in tightly stoppered bottles. The *NH<sub>4</sub> salt* forms needles from H<sub>2</sub>O [α]<sub>D</sub><sup>16</sup> ±20.5° (c 5, H<sub>2</sub>O). [Burgess & Lowry *J Chem Soc* **127** 279 1925, Marsi et al. *J Am Chem Soc* **78** 3063 1956.] The *RS-acid* recrystallises from AcOH. [60g of *(±)-acid* in 60ml of AcOH at 105° gave 40g of crystals has **m 202-203°**]. [Bartlett & Knox *Org Synth* **45** 12 1965, Beilstein **11** IV 642.]

(1*S*)-(+)-Camphor-10-sulfonic acid [3144-16-9] **M 232.3, m 193°(dec), 197-198°**,  $[\alpha]_{546}^{20} +27.5^\circ$  (c 10, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{20} +43.5^\circ$  (c 4.3, EtOH),  $\text{pK}_{\text{Est}} \sim -1$ . Crystallise the acid from ethyl acetate and dry it under vacuum (deliquescent). [Loudon *J Chem Soc* 823 1933, Komppa *J Prakt Chem* 162 19 1943, Beilstein 11 IV 642.] See above for *RS*-isomer.

**Camphor-10-sulfonyl chloride** [1*S*-(+)- 21286-54-4, 1*R*-(-)- 39262-22-1] **M 250.7, m 67-68°, 70°,  $[\alpha]_{\text{D}}^{20}$  (+) and (-) 32.2°** (c 3, CHCl<sub>3</sub>). If free from OH bands in the IR, then recrystallise it from Et<sub>2</sub>O or petroleum ether; otherwise treat it with SOCl<sub>2</sub> at 50° for 30 minutes, evaporate, dry the residue over KOH in a vacuum and recrystallise it. The ( $\pm$ )-acid chloride has **m 85°** [Bartlett & Knox *Org Synth* 45 14 1965]. It is characterised as the *amide* (prisms from EtOH) **m 132°,  $[\alpha]_{\text{D}}^{17}$  (+) and (-) 1.5°** (EtOH). On repeated recrystallisation from EtOH the *anilide* has **m 120.5-121°,  $[\alpha]_{\text{D}}^{25} +76^\circ$**  (c 1, CHCl<sub>3</sub>). [Read & Storey *J Chem Soc* 2761 1930, Sutherland & Shriner *J Am Chem Soc* 58 62 1936, Halterman et al. *J Am Chem Soc* 109 8105 1987, Bartlett & Knox *Org Synth* 45 55 1945, Beilstein 11 IV 650.]

**2,10-Camphorsultam** [1*R*-(+)- 108448-77-7, 1*S*-(-)- 94594-90-8] **M 215.3, m 181-183°, 183-184 185-187°,  $[\alpha]_{\text{D}}^{20}$  (+) and (-) 32°** (c 5, CHCl<sub>3</sub>). The (-)-enantiomer is recrystallised from 95% EtOH and dried in a vacuum desiccator. It dissolves in dilute aqueous NaOH and can be precipitated without hydrolysis by acidifying. It forms the *N*-Na salt in EtOH (by addition of Na to the EtOH solution), and the salt can be methylated with MeI to give the (-)-*N*-Me lactam with **m 80°** after recrystallisation from hot H<sub>2</sub>O, and has  $[\alpha]_{\text{D}}^{25} -59.6^\circ$  (c 5, CHCl<sub>3</sub>) [Shriner et al. *J Am Chem Soc* 60 2794 1938]. [Oppolzer et al. *Helv Chim Acta* 69 1142 1986, Weismiller et al. *Org Synth* 69 154 1955, Beilstein 27 III/IV 1007.]

**4-Carboethoxy-3-methyl-2-cyclohexen-1-one (Hagemann's ester)** [487-51-4] **M 182, b 79-80°/0.2mm, 121-123°/4mm, 142-144°/15mm,  $n_{\text{D}}^{20}$  1.488,  $d_4^{20}$  1.038**. Dissolve the ester in ether, shake with solid K<sub>2</sub>CO<sub>3</sub>, aqueous saturated NaHCO<sub>3</sub>, brine, dry (MgSO<sub>4</sub>) and distil it. The *semicarbazone* has **m 165-167° (169°)**. [Smith & Rouault *J Am Chem Soc* 65 631 1943, Beilstein 10 H 631, 10 I 300, 10 III 2899, 10 IV 2666.]

(-)-Caryophyllene oxide (1-*S*-5*c*-6*t*-epoxy-6*c*,10,10-trimethyl-2-methylene-1*r*,9*t*-bicyclo[7.2.0]undecane) [1139-30-6] **M 220.4, m 62-63°, 63.5-64°, 64°, b 114-117°/1.8mm, 141-142°/11mm,  $d_4^{20}$  0.967,  $n_{\text{D}}^{20}$  1.4956,  $[\alpha]_{\text{D}}^{20} -79^\circ$**  (c 2, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{20} -68^\circ$  (supercooled melt). Purify the oxide by TLC on silica gel with EtOAc/petroleum ether (b 60-80°) (15:85), and recrystallise it from MeOH or \*C<sub>6</sub>H<sub>6</sub>. [NMR: Warnhoff *Can J Chem* 42 1664 1964, Ramage & Whitehead *J Chem Soc* 4336 1954, Beilstein 17 IV 392.]

(+)-Cedrol [octahydro-3,6,8,8-tetramethyl-1-3*a*,7-methanoazulen-6-ol, 8*aS*-6*c*-hydroxy-3*c*,6*t*,8,8-tetramethyl(8*ar*-*H*)-octahydro-3*H*,3*at*,7*t*-methanoazulene], [77-53-2] **m 82-86°, 86-87°,  $[\alpha]_{\text{D}}^{28} +10.5^\circ$**  (c 5, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{18} +13.1^\circ$  (c 5.5, EtOH),  $[\alpha]_{\text{D}}^{18} +14.3^\circ$  (c 10, dioxane). Purify cedrol by recrystallisation from aqueous MeOH. It is estimated colorimetrically with H<sub>3</sub>PO<sub>4</sub> in EtOH followed by vanillin and HCl [Hayward & Seymour *Anal Chem* 20 572 1948]. The 3,5-dinitrobenzoyl derivative has **m 92-93°**. [Stork & Clarke *J Am Chem Soc* 83 3114 1961, Beilstein 6 III 424.]

**Chaulmoogric acid [(13-cyclopent-2-enyl)tridecanoic acid]** [29106-32-9] **M 280.4, m 68.5°, b 247-248°/20mm,  $[\alpha]_{\text{D}}^{20} +60^\circ$**  (c 4, CHCl<sub>3</sub>),  $\text{pK}_{\text{Est}} \sim 5.0$ . Crystallise the acid from petroleum ether or EtOH. The *Me ester* [24828-59-9] has **m 22°, b 227°/20mm** and  $[\alpha]_{\text{D}}^{15} +50^\circ$  (c 5, CHCl<sub>3</sub>). [Mislow & Steinberg *J Am Chem Soc* 77 3807 1955.]

**Chlorendic anhydride (1,4,5,6,7,7-hexachloro-5-norbornene-2,3-dicarboxylic anhydride)** [115-27-5] **M 370.9, m 234-236°. 235-237°, 238°**. Steam distil the anhydride or recrystallise it from H<sub>2</sub>O to yield pure diacid. The pure diacid yields the anhydride with Ac<sub>2</sub>O. [Prill *J Am Chem Soc* 69 62 1947.]

**Chlorocyclohexane (cyclohexyl chloride)** [542-18-7] **M 118.6, b 46-48°/26mm, 142.5°/atm,  $d_4^{20}$  1.00,  $n_{\text{D}}^{25}$  1.46265**. It has been prepared using the general procedure of Norris and coworkers by refluxing the alcohol with 5 volumes of "constant boiling HCl" for 2-4 hours [Norris et al. *J Am Chem Soc* 38 1071 (1077) 1916]. It has also been prepared in 76% yield by the "CaCl<sub>2</sub>" procedure described for chlorocyclopentane (see below

[930-28-9]]. Wash the chlorocyclohexane several times with dilute  $\text{NaHCO}_3$ , then repeatedly with distilled water. Dry it with  $\text{CaCl}_2$  and fractionally distil it slowly at atmospheric pressure or better under vacuum. [Perlman et al. *J Org Chem* **1** 288 (294) 1936, IR: Roberts & Chambers *J Am Chem Soc* **73** 5031 1951, *Beilstein* **5** H 21, **5** I 8, **5** II 11, **5** III 37, **5** IV 48.]

**2-Chlorocyclohexanone** [822-87-7] **M** 132.6, **m** 23°, **b** 82-83°/10mm, 98-99°/14-15mm, **d**<sub>4</sub><sup>25</sup> 1.161, **n**<sub>D</sub><sup>20</sup> 1.484. The chlorocyclohexanone is prepared in a well ventilated fume cupboard by bubbling chlorine (215g, ~3moles, use mercury traps) rapidly through a mixture of cyclohexanone (284g, 3moles) and  $\text{H}_2\text{O}$  (900ml) while being stirred in an ice bath at a reaction temperature between 20° and 50°. After about 45 minutes all the chlorocyclohexanone separates as a heavy oil which is combined with the  $\text{Et}_2\text{O}$  (3 x 150ml) extract of the upper aqueous layer, washed with  $\text{H}_2\text{O}$  (150ml), saturated aqueous  $\text{NaCl}$  (200ml), filtered (by gravity) through anhydrous  $\text{Na}_2\text{SO}_4$ , the solvent is evaporated, the residue is distilled in a Claisen flask, and the fraction boiling below 110° (300-340g, **b** 110°/13mm, 92°/4mm) is collected. This is then carefully fractionated through a heated 42-inch Vigreux column with a variable take-off head to give pure 2-chlorocyclohexanone (240-265g, 61-66%, **b** 90-91°/14-15mm) with 5-13% recovery of cyclohexanone **b** 52°/14-15mm. [Newman et al. *Org Synth Coll Vol* **3** 188 1855, *Beilstein* **7** H 10, **7** I 8, **7** II 11, **7** III 36, **7** IV 32.]

**Chlorocyclopentane (cyclopentyl chloride)** [930-28-9] **M** 104.6, **b** 113-115°/atm, 114°/atm, **d**<sub>4</sub><sup>20</sup> 1.005, **n**<sub>D</sub><sup>20</sup> 1.4512. It is prepared by mixing cyclopentanol (43g) with concentrated  $\text{HCl}$  (125ml) and anhydrous  $\text{CaCl}_2$  (50g), and stirring under reflux on a steam bath for 10 hours. [As  $\text{HCl}$  gas evolves continuously, the reaction should be carried out in an sufficient fume cupboard.] After cooling, the upper layer is collected, washed with brine, saturated aqueous  $\text{NaHCO}_3$  (CARE, as evolution of  $\text{CO}_2$  will occur), brine again and dried over anhydrous  $\text{CaCl}_2$  for at least 24 hours. Filter off the solid and fractionate through an efficient column at atmospheric pressure to obtain a ~ 58% yield of cyclopentyl chloride (30g). Bubbling dry  $\text{HCl}$  gas through the reaction mixture does not appear to increase the yield. See preparation and purification of chlorocyclohexane above. [*Beilstein* **5** IV 18.]

(-)- $\alpha$ -Copaene (1*R*,2*S*,6*S*,7*S*,8*S*-8-isopropyl-1,3-dimethyltricyclo[4.4.0.0<sup>2,7</sup>]dec-3-ene) [3856-25-5] **M** 204.4, **b** 119-120°/10mm, 246-251°, **d**<sub>4</sub><sup>20</sup> 0.908, **n**<sub>D</sub><sup>20</sup> 1.489,  $[\alpha]_D^{20}$  -6.3° (c 1.2,  $\text{CHCl}_3$ ). Purify it by distillation, preferably under vacuum. [Heathcock *J Am Chem Soc* **88** 4110 1966, Heathcock et al. *J Am Chem Soc* **89** 4133 1967, Corey & Watt *J Am Chem Soc* **95** 2303 1973, *Beilstein* **5** IV 1189.]

**Cyclobutane** [287-23-0] **M** 56.1, **m** -50°, -80°, **b** 13°/740mm, 12°/atm, **d**<sub>4</sub><sup>20</sup> 0.721, **n**<sub>D</sub><sup>20</sup> 1.426. This easily liquefiable gas is dried over  $\text{Na}$  at melting ice temperature for 4 days and distilled at low temperature through a Podbielniak precision still. A dry sample has been prepared by passage through  $\text{P}_2\text{O}_5$  and distilled repeatedly until all fractions had similar vapour pressures at 0°. [Cansson & Wat *J Org Chem* **14** 31 1949, Heisig *J Am Chem Soc* **63** 1698 1941, Stodola & Heisig *Org Synth Coll Vol* **III** 213 1955.]

**Cyclobutane carboxylic acid** [3721-95-7] **M** 100.1, **m** 3-4°, -5.4°, **b** 84-84.5°/10mm, 110°/25mm, 135-138°/110mm, 194°/760mm, **d**<sub>4</sub><sup>20</sup> 1.061, **n**<sub>D</sub><sup>20</sup> 1.453, **pK**<sup>25</sup> 4.79. Dissolve the acid in aqueous  $\text{HCO}_3^-$  then acidify with  $\text{HCl}$  and extract it into  $\text{Et}_2\text{O}$ , wash with  $\text{H}_2\text{O}$ , dry ( $\text{Na}_2\text{SO}_4$ ), concentrate to a small volume, then distil it through a glass helices packed column. The *S*-benzylisothiuronium salt has **m** 176° (from  $\text{EtOH}$ ), the anilide has **m** 112.5-113°, and the *p*-toluide has **m** 123°. [Payne & Smith *J Org Chem* **22** 1680 1957, Kantaro & Gunning *J Am Chem Soc* **73** 480 1951, Stodola & Heisig *Org Synth Coll Vol* **III** 213 1955, *Beilstein* **9** H 5.]

(±)-*trans*-Cyclobutane-1,2-dicarboxylic acid [1124-13-6] **M** 144.1, **m** 131°, **pK**<sub>1</sub><sup>25</sup> 4.11 (**pK**<sub>1</sub><sup>20</sup> 3.77), **pK**<sub>2</sub><sup>25</sup> 5.15 (**pK**<sub>2</sub><sup>20</sup> 5.63). Crystallise the acid from  $^*\text{C}_6\text{H}_6$  or  $^*\text{C}_6\text{H}_6/\text{EtOAc}$ . The diphenacyl ester has **m** 98° (from  $\text{EtOH}$ ) and the *p*-bromodiphenacyl ester has **m** 158° (from  $\text{EtOH}$ ). The *cis*-acid isomerises to the *trans*-acid on heating in conc  $\text{HCl}$  at 190°. [Reed *J Chem Soc* 685 1951, Fison et al. *J Am Chem Soc* **51** 1536 1929, Fison et al. *J Am Chem Soc* **56** 1774 1934, **pK**: Bode *Chem Ber* **67** 332 1934, *Beilstein* **9** IV 2788.]

*cis*-Cyclobutane-1,2-dicarboxylic acid [1461-94-5] **M** 144.1, **m** 139.5°, 139-140°, **pK**<sub>1</sub><sup>20</sup> 4.20, **pK**<sub>2</sub><sup>20</sup> 6.56. Purify the acid by crystallisation from  $\text{H}_2\text{O}$  or ligroin, or by hydrolysis of the anhydride [**b** 120-150°/40mm, **m** 77-77.5° (from  $^*\text{C}_6\text{H}_6$ , 74-75° from  $\text{H}_2\text{O}$  or ligroin)] with  $\text{H}_2\text{O}$ . The diphenacyl ester has **m** 113° (from  $\text{EtOH}$ )

and the *p*-bromodiphenacyl ester has **m** 153° (from EtOH/Me<sub>2</sub>CO). [Vogel *Justus Liebigs Ann Chem* **615** 13 1958, Reed *J Chem Soc* 685 1951, Fison et al. *J Am Chem Soc* **56** 1774 1934, pK: Bode *Chem Ber* **67** 332 1934, Beilstein **9** IV 2788.]

**Cyclobutanone** [1191-95-3] **M 70.1, b 96-97°, 99-100°/atm, d<sub>4</sub><sup>20</sup> 0.931, n<sub>D</sub><sup>20</sup> 1.4210, n<sub>D</sub><sup>52</sup> 1.4189.** Treat cyclobutanone with dilute aqueous KMnO<sub>4</sub>, dry it with molecular sieves and fractionally distil it. Purify it *via* the semicarbazone, then regenerate the ketone, dry it (CaSO<sub>4</sub>), and distil it in a stainless steel spinning-band (or Vigreux column). *Alternatively*, purify it by preparative gas chromatography using a Carbowax 20-M column at 80°. (This treatment also removes acetone). It has FT-IR (NaCl) with  $\nu_{\max}$  at 3543.3, 2970.3 (CH), 1783.2 (C=O), 1392.5 (OH enol ?), 1208.9, 1136.4, 1080.8, 724.9, and 461.4 cm<sup>-1</sup>; <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 2.01 (q, 2H, C-3 methylene), 3.09 (t, 4H, C-2 and C-4 methylenes); <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 9.74 (C-3), 47.73 (C-2,4) and 208.90 (C-1). The *oxime* has **m** 84-85° (from petroleum ether) and the *semicarbazone* has **m** 212-212.5° (220-221° from MeOH or H<sub>2</sub>O, Buchanan et al. *J Am Chem Soc* **64** 2701 1942). [Salaun *Org Synth* **57** 36 1977, Fitzer & Quabeck *Synthesis* 299 1987, Beilstein **7** IV 3.]

**Cyclobutylamine** [2516-34-9] **M 71.1, b 82-83°/atm, 83.2-84.2°/760mm, d<sub>4</sub><sup>20</sup> 0.839, n<sub>D</sub><sup>20</sup> 1.437, pK<sup>25</sup> 10.04 (9.34 in 50% aqueous EtOH).** It has been purified by steam distillation. The aqueous distillate (e.g. 2L) is acidified with 3N HCl (90ml) and evaporated to dryness in a vacuum. The *hydrochloride* is treated with a few ml of H<sub>2</sub>O, cooled in ice and a slush of KOH pellets ground in a little H<sub>2</sub>O is added slowly in portions and keeping the solution very cold. The amine separates as an oil from the strongly alkaline solution. The oil is collected, dried over solid KOH and distilled using a vacuum jacketed Vigreux column and protected from CO<sub>2</sub> using a soda lime tube. The fraction boiling at 79-83° is collected, dried over solid KOH for 2 days and redistilled over a few pellets of KOH (**b** 80.5-81.5°). Best distil in a dry N<sub>2</sub> atmosphere. The purity can be checked by GLC using a polyethylene glycol on Teflon column at 72°, 15 psi, flow rate of 102 ml/min of He. The sample can appear homogeneous but because of tailing it is not possible to tell if H<sub>2</sub>O is present. The NMR in CCl<sub>4</sub> should show no signals less than 1 ppm from TMS. The *hydrochloride* has a multiplet at *ca* 1.5-2.6ppm (H 2,2,4,3,3,4,4), a quintet at 3.8 ppm (H 1) and a singlet at 4.75 for NH<sub>2</sub> [Roberts & Chambers *J Am Chem Soc* **73** 2509 1951]. The *benzenesulfonamide* has **m** 85-86° (from aqueous MeOH) and the *benzoyl* derivative has **m** 120.6-121.6°. [Roberts & Mazur *J Am Chem Soc* **73** 2509 1951, Iffland et al. *J Am Chem Soc* **75** 4044 1953, Werner & Casanova Jr *Org Synth Coll Vol V* 273 1973, Beilstein **12** IV 3.]

**Cyclodecanone** [1502-06-3] **M 154.2, m 21-24°, b 100-102°/12mm.** Purify the ketone *via* the *semicarbazone* (**m** 205-207°, from EtOH) and distil it through an efficient column. It sublimes in a vacuum. The *oxime* has **m** 80°, from MeOH or by sublimation in a high vacuum. [Cope et al. *Org Synth Coll Vol IV* 218 1963, Prelog et al. *Helv Chim Acta* **30** 1746 1947, Ruzicka et al. *Helv Chim Acta* **11** 675 1930, Beilstein **7** III 134, **7** IV 76.]

**cis-Cyclodecene** [935-31-9] **M 138.3, m -3°, -1°, b 73°/15mm, 90.3°/33mm, 194-195°/740mm, 197-199°/atm, d<sub>4</sub><sup>20</sup> 0.8770, n<sub>D</sub><sup>20</sup> 1.4854.** Purify it by fractional distillation. It forms an AgNO<sub>3</sub> complex which crystallises from MeOH, **m** 167-187° [Cope et al. *J Am Chem Soc* **77** 1628 1955, IR: Blomqvist et al. *J Am Chem Soc* **74** 3636 1952, Prelog et al. *Helv Chim Acta* **35** 1598 1952].

**cis-cis-trans-1,5,9-Cyclododecatriene (cyclododec-1c,5c,9t-triene)** [2765-29-9] **M 162.3, m -9°, -8°, b 117.5°/2mm, 237-239°/atm, 244°/760mm, d<sub>4</sub><sup>20</sup> 0.907, n<sub>D</sub><sup>20</sup> 1.5129.** Purify the triene by fractional distillation, preferably in a vacuum under N<sub>2</sub>, and it forms an insoluble AgNO<sub>3</sub> complex. [IR: Breil et al. *Makromol Chemie* **69** 28 1963, Beilstein **5** IV 1114.]

**Cyclododecylamine** [1502-03-0] **M 183.3, m 27-29°, b 140-150°/ca 18mm, 280°/atm, pK 9.62 (in 80% methyl cellosolve).** It can be purified *via* the *HCl salt* **m** 274-275° (from EtOH) or the *picrate* **m** 232-234°, and the free base is distilled preferably at water-pump vacuum. It is a strong base and should be stored away from moisture and CO<sub>2</sub>. [Prelog et al. *Helv Chim Acta* **33** 365 1950].

**1,3-Cycloheptadiene** [4054-38-0] **M 94.2, b 55°/75mm, 71.5°/150mm, 120-121°/atm, d<sub>4</sub><sup>20</sup> 0.868, n<sub>D</sub><sup>20</sup> 1.4972.** Purify the diene by dissolving it in Et<sub>2</sub>O, washing with 5% HCl, H<sub>2</sub>O, drying (MgSO<sub>4</sub>), evaporating, and the residue is distilled under dry N<sub>2</sub> through a semi-micro column (some foaming occurs) [Cope et al. *J Am*

*Chem Soc* **79** 6287 1957, UV: Pesch & Friess *J Am Chem Soc* **72** 5756 1950]. [*Beilstein* **5** H 115, **5** III 317, **5** IV 390.]

**Cycloheptane** [291-64-5] **M 98.2, m** ~12°, ~13°, **b 114.4°, 118°/atm**  $d_4^{20}$  **0.812,  $n_D^{20}$  1.4588**. Distil it from sodium using a Vigreux column, under nitrogen. It is highly flammable. [Bocian & Strauss *J Am Chem Soc* **99** 2866 1977, Ruzicka et al. *Helv Chim Acta* **28** 395 1945, *Beilstein* **5** H 92, **5** IV 92.]

**Cycloheptanol** [502-41-0] **M 114.2, m** 2°, **b 77-81°/11mm, 83-84/14mm, 185°/atm**,  $d_4^{20}$  **0.955,  $n_D^{20}$  1.471**. Purify it as described for cyclohexanol. The 2,4-dinitrobenzoyl derivative has **m** 79° and the *allophanate* has **m** 184° (from EtOAc). [Ruzicka et al. *Helv Chim Acta* **28** 395 1945, *Beilstein* **6** H 10.]

**Cycloheptanone (suberone)** [502-42-1] **M 112.2, b 105°/80mm, 172.5°/760,  $d_4^{20}$  0.949,  $n_D^{20}$  1.461**. Shake suberone with aqueous KMnO<sub>4</sub> to remove material absorbing around 230-240nm, then dry it with Linde type 13X molecular sieves and fractionally distil it through a glass helix packed column. [Blicke et al. *J Am Chem Soc* **74** 2924 1952, Dauben et al. *Org Synth Coll Vol IV* 221, 229 1963, *Beilstein* **7** H 13, **7** I 9, **7** II 14, **7** III 46, **7** IV 39.]

**Cycloheptatriene** [544-25-2] **M 92.1, b 60.5°/122mm, 114-115°/atm,  $d_4^{20}$  0.895,  $n_D^{20}$  1.522**. Wash the triene with alkali, then fractionally distil it. Store it under N<sub>2</sub> or Ar as it resinifies in air. [Dryden *J Am Chem Soc* **76** 2814 1954, Kohler et al. *J Am Chem Soc* **61** 1057 1939, *Beilstein* **5** IV 280.]

**Cycloheptylamine** [5452-35-7] **M 113.2, b 50-52°/11mm, 60°/18mm,  $d_4^{20}$  0.887,  $n_D^{20}$  1.472,  $pK_{Est}$  ~10.5 (H<sub>2</sub>O),  $pK^{24}$  9.99 (in 50% aqueous methyl cellosolve)**. It can be purified by conversion to the *hydrochloride* **m** 242-246°, and the free base is distilled under dry N<sub>2</sub> in a vacuum [Cope et al. *J Am Chem Soc* **75** 3212 1953, Prelog et al. *Helv Chim Acta* **33** 365 1950]. [*Beilstein* **12** IV 115.]

**1,3-Cyclohexadiene** [592-57-4] **M 80.1, m** -89°, **b 83-84°/atm,  $d_4^{20}$  0.840,  $n_D^{20}$  1.471**. Distil the diene from NaBH<sub>4</sub> or Na under N<sub>2</sub> and collect it in a trap cooled in Dry Ice. It is highly flammable. [Marvel & Martell, *J Am Chem Soc* **81** 450 1959, *Beilstein* **5** IV 382.]

**1,4-Cyclohexadiene** [628-41-1] **M 80.1, b 83-86°/714mm, 88.3°/741mm, 86-88°/atm, 88.7-89°/760mm,  $d_4^{20}$  0.8573,  $n_D^{20}$  1.4725**. Dry the diene over CaCl<sub>2</sub> and distil it in a vacuum under N<sub>2</sub>. [Hückel & Wörfel *Chem Ber* **88** 338 1955, Giovannini & Wegmüller *Helv Chim Acta* **42** 1142 1959.] [*Beilstein* **5** IV 385.]

**Cyclohexane** [110-82-7] **M 84.2, f 6.6°, b 80.7°,  $d_4^{24}$  0.77410,  $n_D^{20}$  1.42623,  $n_D^{25}$  1.42354**. It is best to purify it by washing with conc H<sub>2</sub>SO<sub>4</sub> until the washings are colourless, followed by water, aqueous Na<sub>2</sub>CO<sub>3</sub> or 5% NaOH, and again water until neutral. It is then dried with P<sub>2</sub>O<sub>5</sub>, Linde type 4A molecular sieves, CaCl<sub>2</sub>, or MgSO<sub>4</sub>, then Na and distilled. Cyclohexane has been refluxed with, and distilled from Na, CaH<sub>2</sub>, LiAlH<sub>4</sub> (which also removes peroxides), sodium/potassium alloy, or P<sub>2</sub>O<sub>5</sub>. Traces of \*benzene can be removed by passage through a column of silica gel that has been freshly heated: this gives material suitable for ultraviolet and infrared spectroscopy. If there is much \*benzene in the cyclohexane, most of it can be removed by a preliminary treatment with nitrating acid (a cold mixture of 30ml conc HNO<sub>3</sub> and 70ml of conc H<sub>2</sub>SO<sub>4</sub>) which converts \*benzene into nitrobenzene. The impure cyclohexane and the nitrating acid are placed in an ice bath and stirred vigorously for 15 minutes, after which the mixture is allowed to warm to 25° during 1 hour. The cyclohexane is decanted, washed several times with 25% NaOH, then water, dried with CaCl<sub>2</sub>, and distilled from sodium. Carbonyl-containing impurities can be removed as described for chloroform. Other purification procedures include passage through columns of activated alumina and repeated crystallisation by partial freezing. Small quantities may be purified by chromatography on a Dowex 710-Chromosorb W gas-liquid chromatographic column. **Flammable liquid**. [Sabatier *Ind Eng Chem* **18** 1005 1926, Schefland & Jacobs *The Handbook of Organic Solvents* (Van Nostrand) p592 1953, *Beilstein* **5** IV 27.]

**Rapid purification:** Distil, discarding the forerun. Stand distillate over Grade I alumina (5% w/v) or 4A molecular sieves.

**Cyclohexane butyric acid** [4441-63-8] **M 170.3, m** 31°, 26.5-28.5°, **b 136-139°/4mm, 169°/20mm,**

**188.8°/46mm, pK<sup>25</sup> 4.95.** Distil the acid through a Vigreux column, and the crystalline distillate is recrystallised from petroleum ether at low temperatures. The *S*-benzylisothiuronium salt has **m** 154-155° (from EtOH) [Friediger & Pedersen *Acta Chem Scand* **9** 1425 1955, English & Dayan *J Am Chem Soc* **72** 4187 1950]. [Beilstein **9** II 15.]

**Cyclohexane carboxylic acid (hexahydrobenzoic acid)** [98-89-5] **M 172.2, m 31-32°, b 63-67°/~0.1mm, 110°/8mm, 232-233°/atm, d<sup>15</sup> 1.480, n<sup>20</sup> 1.460, pK<sup>25</sup> 4.90.** Crystallise the acid from hot H<sub>2</sub>O (solubility is 0.2% w/w at 15°), it is soluble in organic solvents. Also distil it at as high a vacuum as possible and warm the condenser as it solidifies on cooling. The *acid chloride* [2719-27-9] **M 146.6, has b 184°/atm, d<sup>25</sup> 1.096,** the *methyl ester* has **b** 183°/atm, and the *S*-benzylisothiuronium salt has **m** 165-166° (from EtOH). [Beilstein **9** H 7, **9** I 5, **9** II 6, **9** III 15, **9** IV 16.]

**Cyclohexane-1,2-diaminetetraacetic acid H<sub>2</sub>O, (CDTA)** [H<sub>2</sub>O: 12333-90-4; xH<sub>2</sub>O: 13291-61-7] **M 364.4, m >210°(dec), pK<sub>1</sub> 1.34, pK<sub>2</sub> 3.20, pK<sub>3</sub> 5.75 (6.12), pK<sub>4</sub> 9.26 (12.35).** Dissolve CDTA in aqueous NaOH as its disodium salt, then precipitate it by adding HCl. The free acid is filtered off and boiled with distilled water to remove traces of HCl [Bond & Jones *Trans Faraday Soc* **55** 1310 1959]. Recrystallise it from water and dry it *in vacuo*. [Beilstein **13** III 10.]

**cis-Cyclohexane-1,2-dicarboxylic acid (cis-hexahydrophthalic acid)** [610-09-3] **M 172.2, m 191-192°, 191-194°, pK<sub>1</sub><sup>25</sup> 4.25, pK<sub>2</sub><sup>25</sup> 6.74.** It is purified by recrystallisation from EtOH or H<sub>2</sub>O. [Smith & Byrne *J Am Chem Soc* **72** 4406 1950, Abell *J Org Chem* **22** 769 1957, Beilstein **9** III 3812, **9** IV 2801.]

**trans-Cyclohexane-1,2-dicarboxylic acid (trans-hexahydrophthalic acid)** [2305-32-0] **M 172.2, m 227.5-228°, 228-230.5°, pK<sub>1</sub><sup>25</sup> 4.30, pK<sub>2</sub><sup>25</sup> 6.06.** It is purified by recrystallisation from EtOH or H<sub>2</sub>O. It is formed by hydrolysing the anhydride with water. The *dimethyl ester* has **m** 95-96° (from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether). [Abell *J Org Chem* **22** 769 1957, Smith & Byrne *J Am Chem Soc* **72** 4406 1950, Linstead et al. *J Am Chem Soc* **64** 2093 1942, Beilstein **9** III 3812, **9** IV 2801.] The *1R,2R(-)-trans-cyclohexane-1,2-acid* [46022-05-3] has **m** 171-182° and [α]<sub>D</sub><sup>20</sup> -20° (c 1, Me<sub>2</sub>CO).

**cis-Cyclohexane-1,2-dicarboxylic anhydride (cis-hexahydrophthalic anhydride)** [85-42-7, 13149-00-3] **M 154.2, m 32-34°, b 158°/17mm.** It has been obtained by heating the *trans-acid* or *anhydride* at 200°. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>/Et<sub>2</sub>O or distil it. [Kohler & Jansen *J Am Chem Soc* **60** 2145 1938, Abell *J Org Chem* **22** 769 1957, Beilstein **17** II 452, **17** III/IV 5931.]

**trans-Cyclohexane-1,2-dicarboxylic anhydride (trans-hexahydrophthalic anhydride)** [14166-21-3] **M 154.2, m 140-142°, 145-146°.** Crystallise the anhydride from \*C<sub>6</sub>H<sub>6</sub>/Et<sub>2</sub>O. It has been obtained by heating the *cis-acid* or *anhydride* with HCl at 180° for 3 hours. It is formed from the acid by heating in Ac<sub>2</sub>O. It sublimes at 125-135°/0.02mm. [Kohler & Jansen *J Am Chem Soc* **60** 2145 1938, Fichter & Simon *Helv Chim Acta* **17** 1218 1934, Beilstein **9** IV 2802.]

**(±)-trans-1,2-Cyclohexanediol** [1460-57-7] **M 116.2, m 104°, 105°, 120°/14mm.** Crystallise the diol from Me<sub>2</sub>CO and dry it at 50° for several days. It can also be recrystallised from CCl<sub>4</sub> or EtOAc and it can be distilled. The *2,4-dinitrobenzoyl* derivative has **m** 179°. [Winstein & Buckles *J Am Chem Soc* **64** 2780 1942.] [Beilstein **6** IV 5194.]

**trans-1,2-Cyclohexanediol** [1R,2R(-)- 1072-86-2, 1S,2S-(+)- 57794-08-8] **M 116.2, m 107-109°, 109-110.5°, 111-112°, 113-114°, [α]<sub>D</sub><sup>22</sup> (-) and (+) 46.5° (c 1, H<sub>2</sub>O).** The enantiomers have been recrystallised from \*C<sub>6</sub>H<sub>6</sub> or EtOAc. The (±) diol has been resolved *via* the distrychnine salt of the hemisulfate [Hayward et al. *J Chem Soc Perkin Trans 1* 2413 1976], or the *1-menthoxy acetates*. {*l-trans*-diastereoisomeric salt has **m** 64°, [α]<sub>D} -91.7° (c 1.4, EtOH) from petroleum ether or aqueous EtOH and yields the (-)-*trans*-diol} and {*d-trans*-diastereoisomeric salt has **m** 126-127°, [α]<sub>D} -32.7° (c 0.8, EtOH) from petroleum ether or aqueous EtOH and yields the (+)-*trans* diol}. The *bis-4-nitrobenzoate* has **m** 126.5° [α]<sub>D} (-) and (+) 25.5° (c 1.1, CHCl<sub>3</sub>), and the *bis-3,5-dinitrobenzoate* has **m** 160° [α]<sub>D} (-) and (+) 83.0° (c 1.8, CHCl<sub>3</sub>) [Wilson & Read *J Chem Soc* 1269 1935]. [Beilstein **6** III 4060.]</sub></sub></sub></sub>

**cis-1,3-Cyclohexanediol** [823-18-7] **M 116.2, m 86°, 87°, 137°/13mm.** Crystallise the *cis*-diol from ethyl acetate and acetone or distil it in a vacuum. The *dibenzoyl* derivative has **m 65.5°** (from MeOH or petroleum ether). [Rigby *J Chem Soc* 1586 1949, Furberg & Hassell *Acta Chem Scand* 4 518 1950, Beilstein 6 III 4077, 6 IV 5208.]

**trans-1,3-Cyclohexanediol** [5515-64-0] **M 116.2, m 117°, 118-118.5°, 135°/13mm.** Crystallise the *trans*-diol from ethyl acetate or Me<sub>2</sub>CO. The *dibenzoyl* derivative has **m 123.5°** (from EtOH or petroleum ether). [Rigby *J Chem Soc* 1586 1949, Beilstein 6 III 4077, 6 IV 5208.]

**cis-1,4-Cyclohexanediol** [931-71-5] **M 116.2, m 102.5°, 113-114°.** Crystallise the *cis*-diol from acetone (charcoal), then dry and sublime it under vacuum. It also crystallises from Me<sub>2</sub>CO or Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>. The *diacetate* has **m 40.6-41.1°** (from petroleum ether or 34-36° from EtOH). [Grob & Baumann *Helv Chim Acta* 38 604 1955, Owen & Robins *J Chem Soc* 320 1949, Beilstein 6 III 4080, 6 IV 5209.]

**trans-1,4-Cyclohexanediol** [6995-79-5] **M 116.2, m 142.6-143.1°.** Crystallise the *trans*-diol from MeOH or Me<sub>2</sub>CO. The *diacetate* has **m 104.5-105°** (from petroleum ether or 102-103° from EtOH). [Grob & Baumann *Helv Chim Acta* 38 604 1955, Owen & Robins *J Chem Soc* 320 1949, Beilstein 6 III 4080, 6 IV 5209.]

**Cyclohexane-1,3-dione** [504-02-9] **M 112.1, m 107-108°, pK<sub>1</sub><sup>25</sup> 5.25.** Crystallise the dione from \*benzene. Dissolve ~50g of the diol in 140ml of \*C<sub>6</sub>H<sub>6</sub> under N<sub>2</sub>, cool, collect the solid and dry it in a vacuum desiccator overnight. It is unstable and should be stored under N<sub>2</sub> or Ar at ~0°. [Thompson *Org Synth Coll Vol III* 278 1955, Beilstein 7 IV 1985.]

**Cyclohexane-1,4-dione** [637-88-7] **M 112.1, m 76-77°, 78°, 79.5°, 79-80°, b 130-133°/20mm, d<sub>4</sub><sup>91</sup> 1.0861, n<sub>D</sub><sup>102</sup> 1.4576.** Crystallise the dione from water, then \*benzene. It can also be recrystallised from CHCl<sub>3</sub>/petroleum ether or Et<sub>2</sub>O. It has been purified by distillation in a vacuum, and the pale yellow distillate which solidified is then recrystallised from CCl<sub>4</sub> (14.3 g/100 ml) and has **m 77-79°**. The *di-semicarbazone* has **m 231°**, the *dioxime HCl* has **m 150°** (from MeOH/\*C<sub>6</sub>H<sub>6</sub>) and the *bis-2,4-dinitrophenylhydrazone* has **m 240°** (from PhNO<sub>2</sub>). [Nielsen & Carpenter *Org Synth Coll Vol V* 288 1973, IR: LeFevre & LeFevre *J Chem Soc* 3549 1956.] [Beilstein 7 IV 1986.]

**Cyclohexane-1,2-dione dioxime (Nioxime)** [492-99-9] **M 142.2, m 189-190°, pK<sub>1</sub><sup>25</sup> 10.68, pK<sub>2</sub><sup>25</sup> 11.92.** Crystallise Nioxime from alcohol/water and dry it in a vacuum at 40°. Also 2.5g of oxime have been recrystallised from 550ml of H<sub>2</sub>O using Fe free Norit. It forms complexes with Ni and Pd. [Hach et al. *Org Synth* 32 35 1952.] [Beilstein 7 IV 1982.]

**1,4-Cyclohexanedione monoethylene acetal (1,4-dioxaspiro[4.5]decan-8-one)** [4746-97-8] **M 156.2, m 70-73°, 73.5-74.5°.** Recrystallise it from petroleum ether. It sublimes slowly on attempted distillation. Also purify it by dissolving it in Et<sub>2</sub>O and adding petroleum ether (b 60-80°) until turbid, and cooling. [Gardner et al. *J Am Chem Soc* 22 1206 1957, Britten & Lockwood *J Chem Soc Perkin Trans 1* 1824 1974.] [Beilstein 19/4 V 93.]

**cis,cis-1,3,5-Cyclohexane tricarboxylic acid** [16526-68-4] **M 216.2, m 216-218°, pK<sub>Est(1)</sub> ~4.1, pK<sub>Est(2)</sub> ~5.4, pK<sub>Est(3)</sub> ~6.8.** Purify the acid by recrystallisation from toluene/EtOH or H<sub>2</sub>O. It forms a 1.5 hydrate with **m 216-218°**, and a *dihydrate* **m 110°**. Purify it also by conversion to the *triethyl ester* **b 217-218°/10mm, 151°/1mm**, the distillate solidifies on cooling, **m 36-37°**, which is hydrolysed by boiling in aqueous HCl. The *trimethyl ester* can be distilled and recrystallised from Et<sub>2</sub>O, **m 48-49°**. [Newman & Lawrie *J Am Chem Soc* 76 4598 1954, Lukes & Galik *Col Czech Chem Comm* 19 712 1954, Beilstein 9 III 4749.]

**Cyclohexanol** [108-93-0] **M 100.2, m 25.2°, b 161.1°, d 0.946, n<sub>D</sub><sup>20</sup> 1.466, n<sub>D</sub><sup>25</sup> 1.437, n<sub>D</sub><sup>30</sup> 1.462.** Reflux it with freshly ignited CaO, or dry it with Na<sub>2</sub>CO<sub>3</sub>, then fractionally distil it. Redistil it from a very small amount of sodium, about 0.5-2% depending on the amount of H<sub>2</sub>O estimated to be present. It is further purified by fractional crystallisation from the melt in dry air. Peroxides and aldehydes can be removed by prior washing with ferrous sulfate and water, followed by distillation under nitrogen from 2,4-dinitrophenylhydrazine, using a short fractionating column: water distils as the azeotrope. Dry cyclohexanol is *very hygroscopic*, store in a dry

atmosphere. The *3,4-dinitrobenzoate* has **m** 111-112° (EtOH or aqueous EtOH). It has **TOXIC** vapours. [Beilstein 6 III 10, 6 IV 20.]

**Cyclohexanone** [108-94-1] **M** 98.2, **f** -16.4°, **b** 155.7°, **d**<sub>4</sub><sup>20</sup> 0.947, **n**<sub>D</sub><sup>15</sup> 1.452, **n**<sub>D</sub><sup>20</sup> 1.451, **pK**<sup>25</sup> -6.8 (aqueous H<sub>2</sub>SO<sub>4</sub>), **pK**<sup>25</sup> 11.3 (enol), 16.6 (keto). Dry cyclohexanone over MgSO<sub>4</sub>, CaSO<sub>4</sub>, Na<sub>2</sub>SO<sub>4</sub> or Linde type 13X molecular sieves, then distil it. Cyclohexanol and other oxidisable impurities can be removed by treatment with chromic acid or dilute KMnO<sub>4</sub>. More thorough purification is possible by conversion to the bisulfite addition compound, or the semicarbazone, followed by decomposition with Na<sub>2</sub>CO<sub>3</sub> and steam distillation. [For example, equal weights of the bisulfite adduct (crystallised from water) and Na<sub>2</sub>CO<sub>3</sub> are dissolved in hot water and, after steam distillation, the distillate is saturated with NaCl and extracted with Et<sub>2</sub>O which is then dried (anhydrous MgSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>), filtered, and the solvent evaporated prior to further distillation.] **FLAMMABLE**. The *semicarbazone* has **m** 167°, the *4-nitrophenylhydrazone* has **m** 147°, the *2,4-dinitrophenylhydrazone* has **m** 162°, and the *benzal derivative* has **m** 118°. [Beilstein 7 III 14, 7 IV 15.]

**Cyclohexanone-2-carboxylic acid (2-oxocyclohexane carboxylic acid)** [18709-01-8] **M** 142.1, **de-carboxylates at 81-82°**. A preparation which is reproducible on a large scale involves adding cyclohexanone (49g, 52ml, 0.5mole) to a stirred suspension of NaNH<sub>2</sub> in Et<sub>2</sub>O [obtained by dissolving Na (14g, 0.61g atom) in liquid NH<sub>3</sub> (500ml), and then replacing the NH<sub>3</sub> by Et<sub>2</sub>O (500ml) carefully]. The mixture is boiled for 30 minutes, kept at ~35° while dry CO<sub>2</sub> is bubbled through for 3 hours, then poured into a slurry of ice and excess of 2M aqueous HCl, extracted with Et<sub>2</sub>O (5 x 300ml), and the Et<sub>2</sub>O is extracted with excess of saturated aqueous Na<sub>2</sub>CO<sub>3</sub>. The ice cold Na<sub>2</sub>CO<sub>3</sub> solution is acidified with 2M aqueous HCl and again extracted with Et<sub>2</sub>O (5 x 300ml). The combined extracts are dried (MgSO<sub>4</sub>), filtered and evaporated *in vacuo* to give the carboxylic acid (29g, 41%) which can be recrystallised from Et<sub>2</sub>O to form colourless needles. [Christie & Reid *J Chem Soc, Perkin Trans I* 880 1976, Gardner et al. *J Chem Soc* 1910 1764, cf Chiba et al. *Chem Lett* 1387 1978.]

**Cyclohexanone oxime** [100-64-1] **M** 113.2, **m** 91°, **b** 100-105°/10-12mm, 206-210°/atm. Crystallise the oxime from water or petroleum ether (b 60-80°). [Bousquet *Org Synth Coll Vol II* 313 1943, Beilstein 7 III 32, 7 IV 21.]

**Cyclohexanone phenylhydrazone** [946-82-7] **M** 173.3, **m** 77°, 81°. Crystallise it from EtOH.

**Cyclohexene** [110-83-8] **M** 82.2, **M** -104°, **b** 83°, **d**<sub>4</sub><sup>20</sup> 0.810, **n**<sub>D</sub><sup>20</sup> 1.4464, **n**<sub>D</sub><sup>25</sup> 1.4437. Free cyclohexene from peroxides by washing with successive portions of dilute acidified ferrous sulfate, or with NaHSO<sub>3</sub> solution, then with distilled water, drying with CaCl<sub>2</sub> or CaSO<sub>4</sub>, and distilling under N<sub>2</sub>. Alternative methods for removing peroxides include passage through a column of alumina, refluxing with sodium wire or cupric stearate (then distilling from sodium). The diene is removed by refluxing with maleic anhydride before distilling under vacuum. Treatment with 0.1moles of MeMgI in 40ml of diethyl ether removes traces of oxygenated impurities. Other purification procedures include washing with aqueous NaOH, drying and distilling under N<sub>2</sub> through a spinning band column, redistilling from CaH<sub>2</sub>, storing under sodium wire, and passing through a column of alumina, under N<sub>2</sub>, immediately before use. Store it at <0° under argon. [Coleman & Johnstone *Org Synth Coll Vol I* 83 1955, Carson & Ipatieff *Org Synth Coll Vol II* 152 1943, Woon et al. *J Am Chem Soc* 108 7990 1986, Wong et al. *J Am Chem Soc* 109 3428 1987.] [Beilstein 5 IV 218.]

**(±)-2-Cyclohexen-1-ol (3-hydroxycyclohex-1-ene)** [822-67-3] **M** 242.2, **b** 63-65°/12mm, 65-66°/13mm, 67°/15mm, 74°/25mm, 85°/35mm, 166°/atm, **d**<sub>4</sub><sup>20</sup> 0.9865, **n**<sub>D</sub><sup>20</sup> 1.4720. Purify 2-cyclohexen-1-ol by distillation through a short Vigreux column. The *2,4-dinitrobenzoyl* derivative has **m** 120.5°, and the *phenylurethane* has **m** 107°. [Pedersen et al. *Org Synth* 48 18 1968, Cook *J Chem Soc* 1774 1938, Deiding & Hartman *J Am Chem Soc* 75 3725 1953, Beilstein 6 IV 196.]

**Cyclohexene oxide (7-oxabicyclo[4.1.0]heptane)** [286-20-4] **M** 98.2, **b** 131-133°/atm, **d**<sub>4</sub><sup>20</sup> 0.971, **n**<sub>D</sub><sup>20</sup> 1.452.

Fractionate the oxide through an efficient column. The main impurity is probably H<sub>2</sub>O. Dry the oxide over MgSO<sub>4</sub>, filter it, and redistil it several times (**b** 129-134°/760mm). The residue can be hard to remove from the distilling flask. To avoid this difficulty, add a small amount of a mixture of ground NaCl and Celite (1:1) to help break up the residue particularly if hot H<sub>2</sub>O is added. [Osterberg *Org Synth Coll Vol I* 185 1948, *Beilstein* 17 H 21, 17/1 V 203.]

**Cycloheximide (actidione)** [68-81-9] **M** 281.4, **m** 119.5-121°,  $[\alpha]_{546}^{20} +9.5^\circ$  (**c** 2, H<sub>2</sub>O), **pK** 11.2. Crystallise it from water/MeOH (4:1), amyl acetate, isopropyl acetate/isopropyl ether or water. The *acetate* has **m** 150-152° (from aqueous EtOH), the *p-nitrobenzoate* has **m** 215-220°(dec) (from aqueous dioxane) and the *oxime* has **m** 203-204° (from MeOH). [Kornfeld et al. *J Am Chem Soc* 71 155 1949, *Beilstein* 21 IV 6632.]

**Cyclohexylamine** [108-91-8] **M** 99.2, **b** 134.5°,  $d_4^{20}$  0.866,  $d_4^{25}$  0.863,  $n_D^{20}$  1.4593,  $n_D^{25}$  1.456, **pK**<sup>25</sup> 10.63. Dry the amine with CaCl<sub>2</sub> or LiAlH<sub>4</sub>, then distil it from BaO, KOH or Na, under N<sub>2</sub>. Also purify it by conversion to the hydrochloride (which is crystallised several times from water), then liberation of the amine with alkali and fractional distillation under N<sub>2</sub>. The *hydrochloride* has **m** 205-207° (dioxane/EtOH). [Lycan et al. *Org Synth Coll Vol II* 319 1943, *Beilstein* 12 III 10, 12 IV 8.]

**Cyclohexyl bromide** [108-85-0] **M** 156.3, **b** 72°/29mm,  $d_4^{20}$  0.902,  $n_D^{25}$  1.4935. Shake the bromide with 60% aqueous HBr to remove the free alcohol. After removing excess HBr, the sample is dried and fractionally distilled. [IR: Roberts & Chambers *J Am Chem Soc* 73 5031 1951, *Beilstein* 5 III 48, 5 IV 67.]

**1-Cyclohexylethylamine** [*S*-(+)- 17430-98-7, *R*-(-)- 5913-13-3] **M** 127.2, **b** 177-178°/atm,  $d_4^{20}$  0.866,  $n_D^{20}$  1.446,  $[\alpha]_D^{15}$  (-) and (+) 3.2° (**neat**), **pK**<sub>Est</sub> ~10.6. Purify it by conversion to the *bitartrate salt* (**m** 172°), then decomposing with strong alkali and extracting into Et<sub>2</sub>O, drying (KOH), filtering, evaporating and distilling. The *hydrochloride salt* has **m** 242° (from EtOH/Et<sub>2</sub>O),  $[\alpha]_D^{15}$  -5.0° (**c** 10 H<sub>2</sub>O, from (+) amine). The *oxalate salt* has **m** 132° (from H<sub>2</sub>O). The ( $\pm$ )-*base* has **b** 176-178°/760mm, and its *hydrochloride* has **m** 237-238°. [Reihlen et al. *Justus Liebigs Ann Chem* 532 247 1938, *Leithe Chem Ber* 65 660 1932, *Beilstein* 12 III 95.]

**Cyclohexylidene fulvene (6,6-pentamethylene fulvene)** [3141-04-6] **M** 134.2. Purify the fulvene by column chromatography and eluting with *n*-hexane [Abboud et al. *J Am Chem Soc* 109 1334 1987].

**Cyclohexyl mercaptan (cyclohexane thiol)** [1569-69-3] **M** 116.2, **b** 38-39°/12mm, 57°/23mm, 90°/100mm, 157°/763mm,  $d_4^{20}$  0.949,  $n_D^{20}$  1.493, **pK**<sub>Est</sub> ~10.8. Possible impurities are the sulfide and the disulfide. Purify the thiol by conversion to the Na salt by dissolving it in 10% aqueous NaOH, extract the sulfide and disulfide with Et<sub>2</sub>O, and then acidify the aqueous solution (with cooling and under N<sub>2</sub>) with HCl, extract with Et<sub>2</sub>O, dry over MgSO<sub>4</sub>, evaporate and distil it in a vacuum (**b** 41°/12mm). The *sulfide* has **b** 74°/0.2mm,  $n_D^{18.5}$  1.5162 and the *disulfide* has **b** 110-112°/0.2mm,  $n_D^{18.5}$  1.5557. The *Hg-mercaptide* has **m** 77-78° (needles from EtOH). [Naylor *J Chem Soc* 1532 1947, *Beilstein* 6 H 8, 6 I 6, 6 II 14, 6 III 46, 6 IV 72.]

**Cyclohexyl methacrylate** [101-43-9] **M** 168.2, **b** 81-86°/0.1mm,  $d_4^{20}$  0.964,  $n_D^{20}$  1.458. Purify it as for methyl methacrylate (see [80-62-6]). [Tong & Kenyon *J Am Chem Soc* 68 1355 1946, *Beilstein* 6 III 25, 6 IV 39.]

**Cyclononane** [3350-30-9] **M** 140.2, **m** 142.0-142.8°, **b** 220-222°, 100-101.5°/15mm. Purify it *via* the *semicarbazone* (**m** 179.5-180.5° from 90% MeOH) and regenerate it by steam distilling a mixture of 13.1g of semicarbazone, 22g of phthalic anhydride and 45ml of H<sub>2</sub>O. After collecting 300ml of distillate, the latter is extracted with Et<sub>2</sub>O. The dried extract (MgSO<sub>4</sub>) gives on evaporation 8.4g of ketone **b** 100-101.5°/15mm. The *oxime* has **m** 76.5-77.5° (79° from MeOH) and the *iso-oxime* has **m** 138-139°. [Ruzicka et al. *Helv Chim Acta* 32 548 1949.] It has also been repeatedly sublimed at 0.05-0.1mm pressure. [Blomquist et al. *J Am Chem Soc* 74 3639, 3645 1952, *Beilstein* 7 III 110, 7 IV 62.]

**cis,cis-1,3-Cyclooctadiene** [3806-59-5] **M** 108.2, **m** -5°, -49°, **b** 55°/34mm, 142-144°/760mm,  $d_4^{20}$  0.8690,  $n_D^{20}$  1.48921. Purify the diene by GLC. Fractional distillation through a Widmer column gives a mobile liquid, and redistil it with a Claisen flask or through a semi-micro column [Gould, et al. *Anal Chem* 20 361 1948]. **NB:** It has a strong characteristic disagreeable odour detectable at low concentrations and causes headaches on

prolonged exposure. *Do not breath it internally.* [IR: Cope & Estes *J Am Chem Soc* **72** 1128 1950, UV: Cope & Baumgardner *J Am Chem Soc* **78** 2812 1956.] [*Beilstein* **5** IV 401.]

**cis-cis-1,5-cyclooctadiene (COD)** [111-78-4, 1552-12-1] **M 108.2, m -69.5°, -70°, b 51-52°/25mm, 97°/144mm, 150.8°/757mm,  $d_4^{20}$  0.880,  $n_D^{20}$  1.4935.** Purify it by GLC. It has been purified *via* the AgNO<sub>3</sub> salt. This is prepared by shaking with a solution of 50% aqueous AgNO<sub>3</sub> w/w several times (e.g. 3 x 50 ml and 4 x 50 ml) at 70° for *ca* 20 minutes to get a good separation of layers. The upper layers are combined and further extracted with AgNO<sub>3</sub> at 40° (2 x 20 ml). The upper layer (19 ml) of original hydrocarbon mixture gives colourless needles of the AgNO<sub>3</sub> complex on cooling. The adduct is recrystallised from MeOH (and cooling to 0°). The hydrocarbon is recovered by steam distilling the salt. The distillate is extracted with Et<sub>2</sub>O, dried (MgSO<sub>4</sub>), filtered, evaporated and distilled. [Jones *J Chem Soc* 312 1954, [*Beilstein* **5** H 116, **5** IV 403.]

**Cyclooctanone** [502-49-8] **M 126.2, m 42°, 43.8°, b 115-115.5°/60mm.** Purify the ketone by sublimation after drying an ethereal solution over Linde type 13X molecular sieves, filtering and evaporating. The *semicarbazone* has **m** 168-169° (from dioxane) [Kohler et al. *J Am Chem Soc* **61** 1060 1939]. The *oxime* has **m** 36-37° after subliming at high vacuum, or distillation, and has **b** 128-129°/14mm. The *iso-oxime* has **m** 72-73° [Ruzicka et al. *Helv Chim Acta* **32** 548 1949]. [*Beilstein* **7** III 77, **7** IV 49.]

**1,3,5,7-Cyclooctatetraene** [629-20-9] **M 104.2, m -5° to -3°, b 141-141.5°,  $d_4^{20}$  1.537,  $n_D^{25}$  1.5350.** Purify the triene by shaking 3ml with 20ml of 10% aqueous AgNO<sub>3</sub> for 15 minutes, then filtering off the AgNO<sub>3</sub> complex which precipitates. The precipitate is dissolved in water and added to cold concentrated ammonia to regenerate the cyclooctatetraene which is fractionally distilled under vacuum onto molecular sieves and stored at 0°. It is passed through a dry alumina column before use [Broadley et al. *J Chem Soc, Dalton Trans* 373 1986]. [*Beilstein* **5** I 228, **5** IV 1331.]

**cis-Cyclooctene** [931-87-3; 931-88-4] **M 110.2, b 32-34°/12mm, 66.5-67°/60mm, 88°/141mm, 140°/170mm, 143°/760mm,  $d_4^{20}$  0.84843,  $n_D^{20}$  1.4702.** The *cis*-isomer is freed from the *trans*-isomer by fractional distillation through a spinning-band column, followed by preparative gas chromatography on a Dowex 710-Chromosorb W GLC column. It is passed through a short alumina column immediately before use [Collman et al. *J Am Chem Soc* **108** 2588 1986]. It has also been distilled in a dry N<sub>2</sub> glove box from powdered fused NaOH through a Vigreux column, then passed through activated neutral alumina before use [Wong et al. *J Am Chem Soc* **109** 4328 1987]. *Alternatively*, it can be purified *via* the AgNO<sub>3</sub> salt. This salt is obtained from crude cyclooctene (40 ml) by shaking at 70-80° with 50% w/w AgNO<sub>3</sub> (2 x 15 ml) to remove cyclooctadienes (aqueous layer). Extraction is repeated at 40° (4 x 20 ml, of 50% AgNO<sub>3</sub>). Three layers are formed each time. The middle layer contains the AgNO<sub>3</sub> adduct of cyclooctene which crystallises on cooling the layer to room temperature. The adduct (complex 2:1) is highly soluble in MeOH (at least 1g/ml) from which it crystallises in large flat needles when cooled at 0°. It is dried under slight vacuum for 1 week in the presence of CaCl<sub>2</sub> and paraffin wax soaked in cyclooctene. It has **m** 51° and loses hydrocarbon on exposure to air. *cis*-Cyclooctene can be recovered by steam distillation of the salt, collected, dried (CaCl<sub>2</sub>) and distilled in a vacuum. [Braude et al. *J Chem Soc* 4711 1957, AgNO<sub>3</sub>: Jones *J Chem Soc* 1808 1954, Cope & Estes *J Am Chem Soc* **72** 1128 1950, *Beilstein* **5** I 35, **5** IV 263.] **FLAMMABLE LIQUID.**

**cis-Cyclooctene oxide** {(1r, 8c)-9-oxabicyclo[6.1.0]nonane} [286-62-4] **M 126.7, m 56-57°, 57.5-57.8°, 50-60°, b 85-88°/17mm, 82.5°/22mm, 90-93°/37mm, 189-190°/atm.** It can be distilled in a vacuum, and the solidified distillate can be sublimed in a vacuum below 50°. It has a characteristic odour. [IR: Cope et al. *J Am Chem Soc* **74** 5884 1952, *cf trans-isomer*: Cope et al. *J Am Chem Soc* **79** 3905 1957, Reppe et al. *Justus Liebigs Ann Chem* **560** 1 1948].

**Cyclopentadecanone (Exaltone)** [502-72-7] **M 224.4, m 63°, 65°, 65-66°, b 155-157°/5mm.** Subliming Exaltone is better than crystallising it from aqueous EtOH for purification. The *semicarbazone* has **m** 186-187°. [Stevens & Erickson *J Am Chem Soc* **64** 146 1942, Mathur et al. *J Chem Soc* 3505 1963, Biens & Hess *Helv Chim Acta* **71** 1704 1988, *Beilstein* **7** III 203, **7** IV 118.]

**Cyclopentadiene** [542-92-7] **M 66.1, b 41-42°,  $pK^{25}$  15.** Dry the diene with Mg(ClO<sub>4</sub>)<sub>2</sub> and distil it rapidly

as it dimerises readily at room temperature. It should be used immediately or stored in a Dry Ice or an ice-salt bath. **HIGHLY FLAMMABLE**. [Moffett *Org Synth Coll Vol IV* 238 1963.] *Cyclopentadiene Dimer (4,7-methano-3a,4,7,7a-tetrahydroindene)* has [77-73-6], **M** 132.3, **m** 33°, **b** 170°/atm, and **d**<sup>25</sup> 0.986; add ~0.05% of 2,6-di-*tert*-butyl-4-methylphenol as stabiliser. Cyclopentadiene is prepared when required by depolymerising the technical grade dimer by heating (and distilling) it carefully under a fractionating column [Wilkinson *Org Synth Coll Vol IV* 467 1963], as described by Moffett (above reference), or by adding the dimer at a steady rate onto mineral oil heated at 240-270° when it distils off. [Korach et al. *Org Synth* **42** 50 1962]. [Beilstein **5** II 391.]

**Cyclopentane** [287-92-3] **M** 70.1, **b** 49.3°, **d**<sub>4</sub><sup>20</sup> 0.745, **n**<sub>D</sub><sup>20</sup> 1.40645, **n**<sub>D</sub><sup>25</sup> 1.4340. Free it from cyclopentene by two passages through a column of dried and degassed activated silica gel. It occurs in petroleum and is **HIGHLY FLAMMABLE**. [NMR: Christl *Chem Ber* **108** 2781 1975, Whitesides et al. **41** 2882 1976, Beilstein **5** III 10, **5** IV 4.]

**Cyclopentane carbonitrile** [4254-02-8] **M** 95.2, **m** -75.2°, -76°, **b** 43-44°/7mm, 50-62°/10mm, 67-68°/14mm, 74.5-75°/30mm, **d**<sub>4</sub><sup>20</sup> 0.912, **n**<sub>D</sub><sup>20</sup> 1.441. Dissolve the nitrile in Et<sub>2</sub>O, wash it thoroughly with saturated aqueous K<sub>2</sub>CO<sub>3</sub>, dry (MgSO<sub>4</sub>) and distil it through a 10 cm Vigreux column. [McElvain & Stern *J Am Chem Soc* **77** 457 1955, Bailey & Daly *J Am Chem Soc* **81** 5397 1959, Beilstein **9** IV 14.]

**Cyclopentanecarboxaldehyde** [872-53-7] **M** 98.1, **b** 36°/12mm, 74-78°/100mm, 140-141°/atm, **d**<sup>25</sup> 0.919, **n**<sub>D</sub><sup>20</sup> 1.4420-1.4428. Several preparations of this aldehyde have been described but only two will be briefly mentioned here, and both start from cyclohexene. The *first* requires decomposition of the mercuric sulfate complex where, under N<sub>2</sub>, concentrated H<sub>2</sub>SO<sub>4</sub> (43.5ml, 80.0g, 0.82mole) in H<sub>2</sub>O (3L) and HgSO<sub>4</sub> (740.0g, 2.49moles, CARE **POISONOUS**) are stirred to form the deep-yellow basic salt. To this stirred mixture, under N<sub>2</sub> at 55°, is added all at once cyclohexene (101ml, 82.0g, 1mole, freshly distilled b 82-84°/atm [110-83-8]), and the temperature kept at 55-65° for 1 hour (optimal conditions for decomposition of complex). The colour of the mixture turns from deep-yellow to cream. The equipment is altered for distillation at the end of the hour, and the temperature is raised, stirring and the N<sub>2</sub> flow are continued as the crude *aldehyde* and H<sub>2</sub>O distil during *ca* 2 hours. The layers are separated, the H<sub>2</sub>O layer is extracted with Et<sub>2</sub>O (3 x 50ml) which is combined with the product, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, and distilled rapidly (b 74-78°/100mm) to give the *aldehyde* (46-52g, 46-53%), which if not used immediately should be stored in a brown bottle at 0° for under N<sub>2</sub>, after 0.1g of hydroquinone is added as stabiliser. The *aldehyde* readily forms a *trimer*, and does so on prolonged storage. [Grummit et al. *Org Synth Coll Vol 5* 320 1973, English et al. *J Am Chem Soc* **73** 615 1951.] When the distillation residue is cooled a solid may be formed which can be distilled above 78°/100mm to give a clear liquid that solidifies, and can be recrystallised from 95% EtOH to give a white solid **m** 122-124°, identical with *cyclopentanecarboxaldehyde trimer* obtained from the *aldehyde* and 85% H<sub>3</sub>PO<sub>4</sub> [Brook & Wright *Can J Chem* **29** 308 1951].

In the *second* procedure, Tl(III)(NO<sub>3</sub>)<sub>3</sub> · 3H<sub>2</sub>O (4.4g, 10mmol, TTN [13453-38-8]) is dissolved in MeOH (50ml), cyclohexene (10mmol) is added, and the mixture is stirred at room temperature or heated until a starch-iodide paper indicates complete reduction of Tl(III) to Tl(I) (usually within a few minutes). The mixture is filtered, and an alcoholic solution containing 10mmol of 2,4-dinitrophenylhydrazine is added. This is evaporated to 1/3 its volume and after addition of H<sub>2</sub>O (10ml) the mixture is heated on a steam bath for 10 minutes. On cooling to 0°, the *2,4-dinitrophenylhydrazone* [20956-07-4] crystallises out and is recrystallised from EtOH to **m** 195.5-196° (85% yield). Hydrolysis of the hydrazone in the usual way provides the free *aldehyde*. On a preparative scale the precipitated Tl(I)(NO<sub>3</sub>) salt is filtered off, the filtrate is evaporated to a small volume, the mixture of aldehyde and its methyl acetal formed is heated on a steam bath with excess of 5% H<sub>2</sub>SO<sub>4</sub> for 30 minutes, and *cyclopentanecarboxaldehyde* is isolated by ether extraction followed by distillation as in the first preparation above. [Taylor et al. *J Am Chem Soc* **95** 3635 1973.] It should be stabilised with 0.1% of hydroquinone as before.

**Cyclopentane carboxylic acid** [3400-45-1] **M** 114.1, **m** 3-5°, **b** 87-89°/2-3mm, 106.5-107°/10mm, 120-123°/27mm, 216°/atm, **d**<sup>25</sup> 1.053, **n**<sub>D</sub><sup>20</sup> 1.4530, **n**<sub>D</sub><sup>25</sup> 1.4522, **pK**<sup>25</sup> 4.98. If it is discoloured, shake it with saturated aqueous NaCl, extract it with Et<sub>2</sub>O, dry the extract (MgSO<sub>4</sub>), filter, evaporate and distil the residue preferably under a vacuum. The *lachrymatory acid chloride* [4524-93-0] has **M** 132.6, **m** 4° and **b** 161-162°

atm,  $d_4^{20}$  1.091,  $n_D^{20}$  1.4622. [Beilstein 9 H 6, 9 IV 11.] The *amide* [3217-94-5] has  $m$  179°. Like the methyl ester below, the acid can also be prepared *via* a Favorskii reaction, but by using a little over 2 moles of alkoxide (1 mol provides the ester). Thus, to sodium (83g, 3.61 mol) in absolute EtOH (2.3L) is added dropwise 2-chlorocyclohexanone (240.5g, 1.82 mol, [822-87-7]) during several hours, left overnight, the EtOH is distilled off, replaced by an equal volume of H<sub>2</sub>O, acidified to pH 3, and the oily acid that separates is extracted with Et<sub>2</sub>O. The extract is washed with H<sub>2</sub>O, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, and distilled to give the pure *acid* (110g, 53%). [Jackman et al. *J Am Chem Soc* 70 497 1948, Mourisson et al. *Bull Soc Chim Fr* 767 1952.] [Beilstein 6 H 6, 6 I 4, 6 II 6, 6 III 11, 6 IV 9.]

**Methyl cyclopentanecarboxylate**, [4630-80-2] **M 128.1, b 70-73°/48mm, 38-39°/7mm, 158.2°/760mm,  $n_D^{25}$  1.4341**, can be obtained by reacting the acid chloride with MeOH and distilling; but it can also be obtained in ~55% yield directly from 2-chlorocyclohexanone and NaOMe (1 mol) in dry Et<sub>2</sub>O, *via* a Favorskii reaction, and purified by fractional distillation, using a Podbielniak column filled with tantalum wire spirals and a partial reflux head, under reduced pressure [Goheen & Vaughan *Org Synth Coll Vol* 4 594 1963]. The *methyl ester* has also been prepared in 98% yield by direct esterification of cyclopentane carboxylic acid with MeOH in CH<sub>2</sub>Cl<sub>2</sub> in the presence of H<sub>2</sub>SO<sub>4</sub>. The acid has FTIR (CCl<sub>4</sub>) with  $\nu_{\max}$  at 2952(s), 2873 (m), 1734 (s), 1436 (w), 1165 (s) and 1140 (s) cm<sup>-1</sup>; <sup>1</sup>H NMR (75MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 1.5-2.0 (m, 8H, 4-methylenes), 2.7 (pentet, 1H,  $J = 9.0$ Hz), 3.7 (s, 3H, Me); <sup>13</sup>C NMR (300MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 25.7, 29.9, 43.6, 51.5, 177.2, and the GCMS has  $m/z$  at 128 (M<sup>+</sup>) [Davis et al. *J Org Chem* 58 6843 1993].

**1,1-Cyclopentanediactic acid (3,3-tetramethyleneglutaric acid)** [16713-66-9] **M 186.2, m 176-177°, 180-181°,  $pK_1^{25}$  3.80,  $pK_2^{25}$  6.77**. Purify it by recrystallisation from H<sub>2</sub>O and dry it in air, *in vacuo* or over CaCl<sub>2</sub>. However, if it is suspect it is better to convert the acid to the anhydride (8-oxaspiro[4.5]decane-7,8-dione, [5662-95-3], **m 64-66° (68° also reported), b 186°/15mm**) by refluxing it for 7 hours in excess of Ac<sub>2</sub>O, evaporating to dryness, and distilling the residue in a vacuum or recrystallising it from light petroleum. [Kon & Thorpe *J Chem Soc* 115 701 1919.] The anhydride (e.g. 30g) is then hydrolysed by refluxing it for 2.5 hours (i.e. until it dissolves) in aqueous KOH (35g, 3 mols in 55ml of H<sub>2</sub>O), then acidified with concentrated HCl, extracted with Et<sub>2</sub>O (3 x 50ml) after saturating the aqueous solution with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, drying the extract (Na<sub>2</sub>SO<sub>4</sub>), filtering, evaporating the Et<sub>2</sub>O and recrystallising the *diacid* from hot H<sub>2</sub>O. [Vogel *J Chem Soc* 1761 1934.] The *dimethyl ester* [70179-60-3] has **b 141°/17mm** [Vogel *J Chem Soc* 1761 1934, Dickins et al. *J Chem Soc* 1503 1922], and the *diethyl ester* has **b 153°/14mm** [Kon *J Chem Soc* 525 1922]. [Beilstein 9 I 319, and anhydride *Beilstein* 17/11 V 80.]

**Cyclopentane-1,1-dicarboxylic acid** [5802-65-3] **M 158.1, m 183°, 184°, (176-178°),  $pK_1$  3.23,  $pK_2$  4.08**. The dicarboxylic acid can be prepared in 71% yield by using the nickel(0) catalyst 2-{2-[2-(dicyclohexylphosphino)ethyl]pyridine}-4-oxo-2-nickela-3-oxa-cis-bicyclo[3.3.0]octane which promotes the di-carbonylation of cyclopentene (1.16g, 2.44 mmol) in THF (30ml)/pyridine (30ml), in the presence of BeCl (0.2g, 2.50 mmol), by CO<sub>2</sub> (10 bar pressure) in an autoclave, and stirring at 60° for 24 hours. The solvent is evaporated off *in vacuo*, the residue is treated with Et<sub>2</sub>O/HCl and the solid is collected, dried and recrystallised from pentane/Et<sub>2</sub>O (10:1), or H<sub>2</sub>O. The pure *dicarboxylic acid* melts at 183° with decarboxylation to *cyclopentanecarboxylic acid* (see preceding acid), and has FTIR (KBr) with  $\nu_{\max}$  at 3200-2500 (COOH) and 1700 (C=O) cm<sup>-1</sup>; <sup>1</sup>H NMR (200MHz, THF-*d*<sub>8</sub>/TMS) with  $\delta$  at 1.6 (m, 4H, 2-methylenes), 2.1 (m, 4H, 2-methylenes), 9.6 (br, 2H, 2-COOH) ppm; <sup>13</sup>C NMR (50MHz, THF-*d*<sub>8</sub>/TMS) with  $\delta$  at 26.3 (t), 35.2 (t), 60.6 (s), 174.2 ppm; and the MS (70 eV) has  $m/z$  at 158 (M<sup>+</sup>). [Hoberg et al. *Synthesis* 395 1991]. **Dimethyl 1,1-cyclopentanedicarboxylate** [74090-15-6] **M 186.2, b 102°/10mm**, can be prepared in 76% yield from the preceding *methyl cyclopentanecarboxylate* (40 mmol, [4630-80-2]) in THF (60ml), with (iPr)<sub>2</sub>NH (60 mmol) and *n*-BuLi (60 mmol) followed by MeOCOCl (60 mmol), and finally distilling. The *dimethyl ester* has FTIR (CCl<sub>4</sub>) with  $\nu_{\max}$  at 2953(s), 2928 (w), 2876 (w), 1736 (s), 1434 (w), 1266 (s), 1165 (s) and 1159 (m) cm<sup>-1</sup>; <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 1.7 (m, 4H, 2-methylenes), 2.1 (m, 4H, 2-methylenes), 3.7 (s, 6H, 2-Me) ppm; <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 25.3, 34.5, 52.4, 60.2, 173.0 ppm, and the GCMS has  $m/z$  at 187 (M<sup>+</sup>+1) [Davis et al. *J Org Chem* 58 6843 1993].

**1RS,2SR-(meso)-cis-1,2-Cyclopentanedicarboxylic acid** [1461-96-7] **M 158.2, m 139°, 140°, 141°,  $pK_1^{20}$  4.42,  $pK_2^{20}$  6.57**. The *cis*-acid has been prepared by hydrogenation of cyclopent-1-ene-1,2-dicarboxylic acid with Pt as catalyst [Peters *J Chem Soc* 1757 1959], heating diethyl 1-cyanocyclopentane-1,2-dicarboxylate (b

135-136°/3.5mm) with aqueous HCl and isolating the anhydride [Fuson & Cole *J Am Chem Soc* **60** 1237 1938, Dutta *J Indian Chem Soc* **17** 611, 617 1940], but is best prepared by boiling the *trans*-acid with excess of acetic anhydride for at least 2 hours (which provides the *cis*-anhydride), dissolving this in boiling H<sub>2</sub>O or aqueous KOH with cooling (**care**, may be exothermic), and acidifying to pH ~2 in order to precipitate the *cis*-acid. The *cis*-acid crystallises from hot H<sub>2</sub>O as colourless needles melting at 140°, and is rapidly converted to the *cis*-anhydride at 150-160°. The *cis*-acid is also obtained from the *trans*-acid by heating at 300° [Perkin Jnr *J Chem Soc* **51** 247 1887, Perkin Jnr *J Chem Soc* **65** 527 1894]. It is more soluble in H<sub>2</sub>O than the (±)-*trans*-acid. It had not been possible to prepare *trans*-cyclopentane-1,2-dicarboxylic anhydride because all attempts gave the *cis*-anhydride (*cis*-tetrahydro-1*H*-cyclopent[*c*]furan-1,3(3*aH*)-dione, **m** 73.5-74°, **b** 100-102°/1.5mm, [5763-49-5]). [Perkin Jnr *J Chem Soc* **65** 588 1894, cf Fuson & Cole *J Am Chem Soc* **60** 1237 1938, Goldsworthy & Perkin Jnr *J Chem Soc* **105** 2639 1914.]. The nickel(0) catalyst 2-*f*-2-(*dicyclohexylphosphino*)ethylpyridine}-4-oxo-2-nickela-3-oxa-*cis*-bicyclo[3.3.0]octane promotes the carbonylation of cyclopentene in THF by CO at 1bar pressure and -40°, then at ~25°/10 hours followed by treatment with 2N H<sub>2</sub>SO<sub>4</sub> and recrystallisation from pentane/Et<sub>2</sub>O (10:1), which provides cyclopentane-1,2-dicarboxylic acid (78% yield) presumed to be mainly the *cis*-isomer. However, its melting point of 124-125° suggests that it must be contaminated with some of the *trans*-isomer [Hoberg et al. *Synthesis* 395 1991]. Its FT-IR (KBr) has  $\nu_{\max}$  at 3300-2500 (COOH) and 1710 (C=O) cm<sup>-1</sup>; its <sup>1</sup>H NMR (200MHz, THF-*d*<sub>8</sub>/TMS) has  $\delta_{\text{H}}$  at 10.4 (br s, 2H, OH), 2.9 (m, 2H, 2CH), 1.51-2.1 (m, 6H, 3CH<sub>2</sub>), 1.69 (m, 4H); its <sup>13</sup>C NMR (200MHz, THF-*d*<sub>8</sub>/TMS) has  $\delta_{\text{C}}$  at 176.10 (s, carboxylic C), 47.5 (d, C1 and C2), 29.4 (t, C-3 and C-5) and 24.4 (t, C-4); and its MS (70 eV) has *m/z* = 158 (M<sup>+</sup>), see GC-MS [below, Lu et al. *J Org Chem* **55** 2503, 2507 1990]. **Dimethyl *cis*-cyclopentane-1,2-dicarboxylate**, **b** 129°/20mm, **n<sub>D</sub><sup>21</sup>** 1.4528, [4841-91-2], is obtained by treating the acid with diazomethane in Et<sub>2</sub>O, evaporating and distilling the residue [Owen & Peto *J Chem Soc* 2383, 2386 1955]. The ***cis*-imide** (*cis*-tetrahydrocyclopent[*c*]pyrrole-1,3(2*H*,3*aH*)-dione, [5763-44-0]) has **m** 85-87° (also 90° reported) after recrystallisation from H<sub>2</sub>O or petroleum ether (b 40-60°) [Menon & Simonsen *J Chem Soc* 302 1929, Rice et al. *J Org Chem* **24** 7 1959], and ***trans*-cyclopentane-1,2-dicarboxylic acid dihydrazide**, prepared from the dimethyl ester and NH<sub>2</sub>NH<sub>2</sub> · H<sub>2</sub>O in EtOH and recrystallised from MeOH/Et<sub>2</sub>O, has **m** 125-126° [Müller et al. *J Am Chem Soc* **73** 2489 1951]. [Beilstein **9** H 728, **9** III 3807, **9** IV 3793.]

***1*RS*,2*RS**(±)-*trans*-1,2-Cyclopentanedicarboxylic acid** [1461-97-8] **M** 158.2, **m** 160-161°, 161.5°, 162°, 162-163°, 163°, 163-165°, 164-165°, **pK<sub>1</sub><sup>20</sup>** 4.14, **pK<sub>2</sub><sup>20</sup>** 5.99. Several syntheses of the (±)-*trans*-acid have been described. The first preparation involved tetraethyl cyclopentane-1,2,3,4-tetracarboxylate [b 185-190°/1mm, from trimethylene-1,3-dibromide (96g), Na (22g) and diethyl malonate (150g) in absolute EtOH (300ml)] which was refluxed with 2-3 volumes of AcOH, 1 volume of concentrated H<sub>2</sub>SO<sub>4</sub> and 1 volume of H<sub>2</sub>O for 2 days, and the AcOH and H<sub>2</sub>O were distilled off to give *trans*-cyclopentane-1,2-dicarboxylic acid [Perkin Jnr *J Chem Soc* **65** 587 1894, Perkin Jnr *J Chem Soc* **51** 224 1887, Wassermann *Helv Chim Acta* **13** 207 1930, Perkin Jnr & Prentice *J Chem Soc* **59** 828 1891]. Other preparations include heating the *cis*-acid with concentrated HCl at 180° [Perkin Jnr *J Chem Soc* **65** 577, 590 1894], refluxing diethyl 1-cyanocyclopentane-1,2-dicarboxylate with concentrated aqueous HCl for 4 days which provided a 92% yield of the *trans*-acid [Fuson et al. *J Org Chem* **10** 121, 126 1945, Fuson & Cole *J Am Chem Soc* **60** 1237 1938], and electrochemical coupling between 1,3-dibromopropane and dimethyl maleate which gave a 10/1 ratio of *trans/cis* dimethyl esters in 42% yield [Lu et al. *J Org Chem* **55** 2503, 2507 1990]. Invariably the crude *trans*-diesters formed are contaminated with some *cis*-isomer and it was found that hydrolysing the ester mixture by refluxing the acidic solution for 4 to 5 days provided only the *trans*-acid in ~95% yields. The *trans*-acid crystallises (charcoal) from hot H<sub>2</sub>O, and it is easily soluble in AcOH, and EtOH, but very soluble in Et<sub>2</sub>O, CHCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub> and petroleum ether. Its FT-IR (Nujol) has  $\nu_{\max}$  at 3008.3, 2653.8, 1692.6, 1419.2, 1311.1, 1253.5, 1199.3, 935.7 and 689.4 cm<sup>-1</sup>; its <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 12.2 (br s, 2H, OH), 2.96 (m, 2H, C1 and C2 methines), 1.99 (m, 2H, ), 1.69 (m, 4H); its <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 176.10 (s, carboxylic C), 46.77 (s, C1 and C2), 30.03 (s, C3 and C5) and 25.01 (s, C-4); and its GC-MS has *m/z* (relative intensity) *trans*-isomer: 157(37, M<sup>+</sup>-2), 155(15), 126(67), 113(29), and 67(100), and *cis*-isomer: 157(6, M<sup>+</sup>-2), 155(40), 126(32), 113(56), and 67(100) [Lu et al. *J Org Chem* **55** 2503, 2507 1990]. **Dimethyl *trans*-cyclopentane-1,2-dicarboxylate**, **b** 118.5-119°/17mm, **n<sub>D</sub><sup>20</sup>** 1.4491, [941-75-3], is obtained by treating the acid with diazomethane in Et<sub>2</sub>O, evaporating and distilling the residue [Owen & Peto *J Chem Soc* 2383, 2386 1955]. ***trans*-cyclopentane-1,2-dicarbonyl dichloride**, has **b** 98-100°/2mm [Aspinall & Baker *J Chem Soc* 743, 747 1950], and ***trans*-cyclopentane-1,2-dicarboxylic acid**

*dihydrazide*, prepared from the dimethyl ester and  $\text{NH}_2\text{NH}_2 \cdot \text{H}_2\text{O}$  in EtOH, and recrystallised from EtOH, has **m 224-225°** [Müller et al. *J Am Chem Soc* **73** 2489 1951]. [Beilstein **9** H 728, **9** I 316, **9** III 3807, **9** IV 2793.]

**1S,2S-(+)-trans-1,2-Cyclopentanedicarboxylic acid** [21917-20-4] **M 158.2, m 181°**,  $[\alpha]_{\text{D}} +87.6^\circ$  (**c 0.88, H<sub>2</sub>O**) and **1R,2R(-)-trans-1,2-cyclopentanedicarboxylic acid** [17224-73-6] **m 180-181°**,  $[\alpha]_{\text{D}} -85.9^\circ$  (**c 1.2, H<sub>2</sub>O**). Optical resolution of the ( $\pm$ )-acid (above) is readily achieved *via* separation of the diastereoisomeric brucine salts (the base being *l*-brucine). Thus pure ( $\pm$ )-acid (15g, ~0.1mole) is dissolved in hot  $\text{H}_2\text{O}$ , and while stirring, the finely powdered brucine tetrahydrate (90g, 0.2 moles, [5892-11-5], care **POISONOUS**) that is added gradually dissolves. At the end of the addition, the mixture is stirred for ~10 minutes and filtered hot from a small amount of undissolved brucine. The brucine is washed well with hot  $\text{H}_2\text{O}$ , and the combined filtrate and washings are concentrated on a boiling water bath until crystals commence to form on the surface. When the liquid is cooled and vigorously stirred, copious crystals of the *d*-acid (*l*-base)<sub>2</sub> separate, and the whole mixture sets into a mass of crystals. These are filtered off and recrystallised from hot  $\text{H}_2\text{O}$  to constant specific rotation. An analytical sample of the *d*-acid (*l*-base)<sub>2</sub> is obtained by dissolving it in hot  $\text{H}_2\text{O}$  and allowing it to crystallise slowly in a desiccator over concentrated  $\text{H}_2\text{SO}_4$ . The decahydrate is obtained on drying the crystalline plates in air but they effloresce in a vacuum desiccator or on heating at 125° for 1 hour to give *anhydrous d-acid (l-base)<sub>2</sub>* that has  $[\alpha]_{\text{D}} -19.9^\circ$  (**c 1.3, H<sub>2</sub>O**). The *d*-acid (*l*-base)<sub>2</sub> is then decomposed by dissolving it in hot  $\text{H}_2\text{O}$ , adding aqueous ammonia, filtering off the solid brucine, washing it well with aqueous ammonia, and the combined hot filtrate and washings are cooled and acidified with hydrochloric acid. On cooling the (+)-acid (m 178-180°) separates and recrystallises from hot  $\text{H}_2\text{O}$  as plates with m 180°. **Diethyl 1S,2S-(+)-trans-1,2-cyclopentanedicarboxylate** **b 170°/100mm**,  $[\alpha]_{\text{D}} +70.3^\circ$  (**c 1.3, Me<sub>2</sub>CO**), is prepared by boiling the (+)-acid with 5 times its weight of 10% ethanolic  $\text{H}_2\text{SO}_4$  for 6 hours, cooling, adding  $\text{H}_2\text{O}$ , the ester is extracted into  $\text{Et}_2\text{O}$  which is washed with dilute  $\text{Na}_2\text{CO}_3$ , dried ( $\text{Na}_2\text{SO}_4$ ), filtered, evaporated and the residue is distilled in a vacuum. The **1S,2S-(+)-trans-dianilide**, **m 245-247°**,  $[\alpha]_{\text{D}} +110.1^\circ$  (**c 0.85, Me<sub>2</sub>CO**) is prepared by heating the (+)-acid with  $\text{SOCl}_2$  in a sealed tube at 100° for 1 hour, evaporating, the di-acid chloride formed is then dissolved in  $\text{C}_6\text{H}_6$ , treated with excess of aniline, evaporated again, the residue is stirred with dilute aqueous HCl, the precipitate is collected and recrystallised twice from large volumes of MeOH.

**1R,2R(-)-trans-1,2-cyclopentanedicarboxylic acid** is obtained by concentrating the mother liquors from the first crystallisation of the *d*-acid (*l*-base)<sub>2</sub> above, which provides copious crystals of the more soluble **diastereoisomeric l-acid (l-base)<sub>2</sub>**. Treatment of this salt with aqueous ammonia as above followed by recrystallisation from  $\text{H}_2\text{O}$  gives the pure (-)-acid with m 180-181°. **Diethyl 1R,2R(-)-trans-1,2-cyclopentanedicarboxylate**, prepared in the same way as its enantiomer above has similar properties except for the optical rotation which is  $[\alpha]_{\text{D}} -69.8^\circ$  (**c 1.7, Me<sub>2</sub>CO**). [Goldsworthy & Perkin Jr *J Chem Soc* **105** 2639 1914]. [Beilstein **9** I 316] The absolute configuration of the (-)-acid was determined as **1R,2R** by comparing the ORD spectra of the respective 1,2-bis(N-methylthionamide) [1,2-(CSNHMe)<sub>2</sub>] with those of the corresponding 1,2-disubstituted cyclopropane, cyclobutane and cyclohexane homologues [Inoue et al. *Tetrahedron* **23** 3237 1967]. [Beilstein **9** III 3807.]

**1RS,2SR-cis-1,2-Cyclopentanediol (meso-cis-cyclopentane-1,2-diol)** [5057-98-7] **M 102.1, m 24-28°, 30°, b 88-92°/2mm, 108-109°/20mm, 123.5°/29mm, n<sub>D</sub><sup>21</sup> 1.4770**. The *cis*-diol can be obtained by treating cyclopentene (21.8g, [142-29-0]) in EtOH (600ml) at -40° with a solution of  $\text{KMnO}_4$  (40g) and anhydrous  $\text{MgSO}_4$  (30g) in  $\text{H}_2\text{O}$  (800ml) dropwise, with vigorous stirring, during 2 hours. The  $\text{MnO}_2$  is filtered off, washed with hot  $\text{H}_2\text{O}$ , the combined aqueous filtrates are evaporated *in vacuo* to 100ml and continuously extracted with  $\text{Et}_2\text{O}$  (liquid-liquid extractor) for 48 hours. The dried ( $\text{MgSO}_4$ )  $\text{Et}_2\text{O}$  extract is evaporated and the residual oil is distilled to give the *cis*-diol (11.1g, 34%). The *cis-di-p-nitrobenzoate* has **m 117° (117-118°)** (from EtOH) the *cis-di-methanesulfonate* has **m 82°** (needles from MeOH), and the *cis-di-toluene-p-sulfonate* has **m 92°** (prisms from EtOH) are obtained in the usual way in pyridine. The *cis-monoacetate* has **b 104-106°/20mm, n<sub>D</sub><sup>17</sup> 1.4576**, the *cis-monoacetate-mono-p-nitrobenzoate* has **m 93-94°** (nodules from EtOH), and the *cis-diacetate* has **m -5°, b 80°/2mm**. The *cis-dibenzoate* has **m 47°**. [Owen & Smith *J Chem Soc* 4026 1952, Beilstein **6** I 369, **6** II 742, **6** III 4053, **6** IV 5187.]

**1RS,2RS-(±)-trans-1,2-Cyclopentanediol [trans-(±)-cyclopentane-1,2-diol]** [5057-99-8, (±) 86703-52-8] **M 102.1, m 50°, 55°, b 93°/2mm, 122°/10mm, 136°/22mm**. The *cis*-diol can be prepared in a seven step synthesis from *racemic* diethyl tartrate [Cunningham & Kündig *J Org Chem* **53** 1823 1988], from *trans*-2-tri-

methylsilylcyclopentan-1-ol and MeOH [Kono & Ngai *Org Prep Proced Int* **6** 19 1974], and by performate oxidation of cyclopentene which is described herein. A mixture of 30% aqueous H<sub>2</sub>O<sub>2</sub> (13g) and formic acid (105g) is added to cyclopentene (6.6g, b 44-45°/atm [142-29-0]) under a reflux condenser, when much heat is generated, the temperature is allowed to fall to 40° and is so maintained for 4 hours. The solvent is removed *in vacuo*, the residue is dissolved in 10% aqueous NaOH (50ml), boiled under reflux for 45 minutes, the solvent is distilled off, and the residue is twice distilled under reduced pressure to give pure *trans*-pentane-1,2-diol (6g, 60%) which solidified in the distillate. The *trans*-dibenzoate has **m 63°**, and the *trans*-(±)-*di-p*-nitrobenzoate (**m 145°**, from EtOH), the *trans*-(±)-*di*-methanesulfonate (**m 82°**, small needles from EtOH), and the *trans*-(±)-*di-p*-toluenesulfonate (**m 109°**, colourless plates from MeOH) are prepared in the usual way in pyridine. [Owen & Smith *J Chem Soc* 4026 1952, Englund *J Prakt Chem* **129** 1, 10 1931, Beilstein **6** H 739, **6** I 369, **6** II 743, **6** III 4053, **6** IV 5187.]

**1*R*,2*R*-(−)-*trans*-1,2-Cyclopentanediol** [*trans*-(−)-cyclopentane-1,2-diol] [(−) 930-46-1, (−)86703-52-8] **M 102.1**, **m 48-48.5°**, [α]<sub>D</sub><sup>20</sup> −24.88° (c 6, EtOH) and **1*S*,2*S*-*trans*-1,2-cyclopentanediol** [*trans*-(+)-cyclopentane-1,2-diol] [(+)107492-82-0], **m 47-48°**, [α]<sub>D</sub><sup>20</sup> +24.54° (c 5.4, EtOH). The *1*R*,2*R**-enantiomer is obtained optically pure in a seven step synthesis from diethyl D(-)-tartrate by benzylation (TiOEt/PhCH<sub>2</sub>Br, 81%), reduction (LAH/Et<sub>2</sub>O, 94%) to *2*R*,3*R*-(−)-2,3-bis(benzyloxy)-1,4-butanediol*, ditosylation (TsCl/py, 87%), dibromination (with LiBr/DMSO, 94%) to *2*R*,3*R*-(−)-2,3-bis(benzyloxy)-1,4-butanedibromide*, formation of the cyclopentane ring (PhSO<sub>2</sub>CH<sub>2</sub>SO<sub>2</sub>Ph/K<sub>2</sub>CO<sub>3</sub>/DMF, 92%) to give *1*R*,2*R*-(−)-1,2-bis(benzyloxy)-4,4-bis(phenylsulfonyl)cyclopentane*. Removal of the bis-sulfonyl groups (Mg/MeOH, 82%) provide *1*R*,2*R*-(−)-1,2-bis(benzyloxy)cyclopentane*, and finally debenylation (H<sub>2</sub>/10% Pd/C, 99%) provides pure *1*R*,2*R*-(−)-cyclopentane-1,2-diol* in 47% overall yield. The final step only is detailed here. The mixture of *trans*-(−)-1,2-bisbenzyloxycyclopentane (5.11g, 18.1mmol) in EtOH (35ml) containing 10% Pd/C (1.35g) is degassed three times, H<sub>2</sub> gas is introduced into the flask at a pressure of 6 atmospheres for 5 hours, the catalyst is filtered off, and the filtrate is evaporated to provide the *1*R*,2*R*-*trans*-(−)-diol* (8.2g, 99%) as a white hygroscopic solid **m 47-48°** which should be stored under dry N<sub>2</sub> in a sealed vessel. The enantiomeric *1*S*,2*S*-*trans*-(−)-diol* can be prepared in 40% overall yield from diethyl L(+)-tartrate *via* the same high yielding steps. Its IR (CH<sub>2</sub>Cl<sub>2</sub>) has ν<sub>max</sub> at 3600, 3390br, 2920, 1450, 1280, 1075, 1040 and 970 cm<sup>-1</sup>; and its <sup>1</sup>H NMR (200MHz, CDCl<sub>3</sub>, TMS) with δ at 3.93-4.06 (m, 2H), 2.24 (s, 2H, exchanges with D<sub>2</sub>O), 1.86-2.12 (m, 2H), 1.63-1.80 (m, 2H) and 1.42-1.62 (m, 2H). The enantiomers are >96% enantiomerically pure. The **1*R*,2*R*-(−)** [113350-86-0] and **1*S*,2*S*-(+)** [113303-27-8] **1,2-bis(dimethoxyphosphinoxy)cyclopentane** derivatives, prepared from the respective diols (2Net<sub>3</sub>/2(MeO)<sub>2</sub>PCl, in 80% and 73% yields respectively), are yellow oils with **b 120°/0.1mm** and [α]<sub>D</sub><sup>20</sup> −41.39° and +41.78° (c 4.8, CH<sub>2</sub>Cl<sub>2</sub>). The **1*R*,2*R*-(−)** [113303-35-8] and **1*S*,2*S*-(+)** [113303-28-9] **1,2-bis(difluorophosphinoxy)cyclopentane** derivatives, prepared from the respective diols (PCl<sub>3</sub> in CH<sub>2</sub>Cl<sub>2</sub>/SbF<sub>3</sub> in pentane, 49% and 48% respectively) are mobile oils with **b 58°/80mm** and [α]<sub>D</sub><sup>20</sup> −12.22° and +12.96° (c 4.8, CH<sub>2</sub>Cl<sub>2</sub>). [Cunningham & Kündig *J Org Chem* **53** 1823 1988].

**1,2-Cyclopentanedione** [3008-40-0] **M 98.1**, **m 55-56°**, **56°**, **b 67-69°/1.4mm**, **87-88°/16mm**, **97°/20mm**, **105°/20mm**, **pK<sup>20</sup> 9.14 (H<sub>2</sub>O)**. The dione has a weak amine-like odour and is soluble in H<sub>2</sub>O, EtOH, and Et<sub>2</sub>O, but poorly soluble in petroleum ether and CS<sub>2</sub>, and is steam volatile. It has been prepared by various routes from cyclopentanone *via* oxidation by SeO<sub>2</sub> (23%) [Goto et al. *Bull Chem Soc, Jpn* **52** 2589 1952], conversion to the α-methylthio ether then oxidation with CuCl<sub>2</sub>/CuO (42%) [Gregoire et al. *J Org Chem* **51** 1419 1986], and autooxidation with O<sub>2</sub>/*t*-BuOK/DME-*t*-BuOH (35.7%). Only the first is described because of its simplicity. *The following procedure should be carried out in a well-ventilated fume cupboard as selenium compounds have foul odours.* To cyclopentanone (1.5L) warmed at 30° is added dropwise, with stirring, a solution of selenium dioxide (361g, 3.25ml) in dioxane (420ml) and H<sub>2</sub>O (155ml) during 24 hours, and stirring is continued for a further 18 hours at 35°. The red selenium that is formed is filtered off (Buchner) and extracted with EtOH (500ml) by boiling for 3 hours. The liquids are combined and distilled through a 60cm Vigreux column at 20mm, the lower boiling portion (b 35-90°) is removed, and the remainder is distilled at 10mm until a thick brown residue is left in the distilling flask. This distillate is redistilled, and the fraction with b 86-88°/16mm is collected to give pure *1,2-dione* (74g, 23% based on SeO<sub>2</sub> used) which solidifies (m 56-58°). It should be stored at −30° under N<sub>2</sub>, in the dark. Cyclopentane-1,2-dione has IR (NaCl) with ν<sub>max</sub> at 1650 (C=C), 1700 (C=O), 3300 (OH enol) cm<sup>-1</sup>; <sup>1</sup>H NMR (400MHz, CCl<sub>4</sub>, TMS) with δ at 2.20-2.67 (m, 4H, 2-methylenes), 6.47-6.62 (m, 1H, C=CH), 6.93-7.36 (m, 1H, C=COH); <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>, TMS) with δ 22.01, 32.58, 131.36, 153.62, 187.53.

In ethanolic solution the 'dione' is almost entirely in the *mono-enol*-form which slowly tautomerises to a mixture containing ~9% of the *diketo*-form after 36 hours at 50°. The UV in hexane has  $\lambda_{\max}$  (log  $\epsilon$ ) at 246nm (4.003) and 300nm (1.900) due mostly to the *diketo*-form, whereas in 0.05N aqueous NaOH it has  $\lambda_{\max}$  (log  $\epsilon$ ) at 288nm (3.383) due to the *enol*-form [Hesse & Krebbiel *Justus Liebigs Ann Chem* **563** 35 1934].

*Cyclopentane-1,2-dione dioxime* (see above [6635-29-6]) decomposes at ~210° after recrystallisation from H<sub>2</sub>O, and a 0.01M aqueous solution of it is used for the gravimetric determination of Ni. [Voter et al. *Anal Chem* **21** 1320 1949.] The *bis-phenylhydrazone* has **m 146°**.

**1,3-Cyclopentanedione** [3859-41-4] **M 98.1, m 149-150°, 151-152.5°, 151-154°, 151-153°, pKa 4.5.** Purify the dione by Soxhlet extraction with CHCl<sub>3</sub>. The CHCl<sub>3</sub> is evaporated and the residue is recrystallised from EtOAc and/or sublimed at 120°/4mm. [IR: Boothe et al. *J Am Chem Soc* **75** 1732 1953, DePuy & Zaweski *J Am Chem Soc* **81** 4920 1959, Beilstein **7** IV 1981.]

**Cyclopentanol** [96-41-3] **M 86.1, m -20°, -19.5°, b 53°/10mm, 56.9-57.4°/34mm, 139.5°/752mm, 140.85°/760mm, d<sub>4</sub><sup>20</sup> 0.9478, n<sub>D</sub><sup>20</sup> 1.4531.** Cyclopentanol has been prepared by hydration of cyclopentene, or from cyclopentanone (see next entry) by catalytic reduction with H<sub>2</sub> (Ni/MeOH, Ra-Ni/EtOH, PtO<sub>2</sub>, Pt-black in EtOH or AcOH, Cu-chromite), Al(isoPrO)<sub>3</sub>/NaOH, LAH/Et<sub>2</sub>O or NaBH<sub>4</sub> in MeOH or H<sub>2</sub>O. The last named has many advantages in ease of use and efficiency. Thus cyclohexanone in H<sub>2</sub>O (solubility is 15% at 10°, 8.7% at 20°, 5% 30°) is stirred with excess of NaBH<sub>4</sub> (theoretically 0.25 mol/mol ketone) at room temperature until the reaction is complete (addition of a drop of dilute HCl causes effervescence). The pH of the solution is adjusted to ~3-4, NaCl is added to almost saturation and the alcohol is extracted several times with Et<sub>2</sub>O, dried (K<sub>2</sub>CO<sub>3</sub>), filtered, and the filtrate is distilled to give cyclopentanol in ~90% yield. [Chaikin & Brown *J Am Chem Soc* **71** 122 1949]. Its FT-IR (neat) has  $\nu_{\max}$  at 3334.9 (br OH), 2959.6 (CH), 1437.2 (OH ?), 1341.3, 1282.9, 1174.3, 1073.7, 994.3, and 837.1 cm<sup>-1</sup>; the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 1.55 (m, 4H, C-3 and C-4 -methylenes), 1.75 (m, 4H, C-2 and C-5 -methylenes), 2.30 (s, 1H, H-1) and 4.30 (brs, 1H, OH); and its <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 23.32 (C-3,4), 35.51 (C-2,5) and 73.86 (C-1). At atmospheric pressure cyclopentanol forms azeotropes with H<sub>2</sub>O (96.3°, 42%), tetrachloroethylene (118.8°, 8%), chlorobenzene, *m*-xylene (132.8°, 40%), *p*-xylene (132.8°, 38%), *n*-Bu<sub>2</sub>O (139.0°, 75%), among other solvents. It can be characterised as the *4-nitrobenzoate* (**m 56°**), *3,5-dinitrobenzoate* (**m 123°**), and *phenylcarbamate* (**m 137-138°**). [Beilstein **6** H 5, **6** I 3, **6** II 3, **6** III 4, **6** IV 5.]

**Cyclopentanone (dumasine, adipic ketone)** [*keto-form* 120-92-3, *enol-form* 59557-02-7] **M 84.1, m -58°, b 130°/atm, d<sub>4</sub><sup>20</sup> 0.947, n<sub>D</sub><sup>20</sup> 1.4366, n<sub>D</sub><sup>25</sup> 1.4340. pK<sup>25</sup> 16.7.** This is a flammable liquid (flash point 26-31°) with a pleasant odour, but which is a strong **SKIN** and **EYE IRRITANT**, and should be used in a well ventilated fume cupboard. It is sparingly soluble in H<sub>2</sub>O, and forms explosive mixtures with HNO<sub>3</sub> and H<sub>2</sub>O<sub>2</sub>. It is volatile with Et<sub>2</sub>O and care should be taken that careful fractionation is required when it is extracted with Et<sub>2</sub>O. It is prepared by heating, in a distilling flask, a powdered mixture of adipic acid (200g, 1.34moles) and ground Ba(OH)<sub>2</sub> (10g) in a metal bath (e.g. Wood's metal) to 285-295° during 1-1.5 hours. [The temperature is best controlled with the thermometer within 5mm above the bottom of the flask and the temperature held as near as 290° as possible to minimising the amount of adipic acid which distils off.] During the heating period, the *cyclopentanone* distils slowly with small quantities of adipic acid, and after a further 2 hours of heating a very small amount of dry residue remains in the flask. The distillate is separated from some H<sub>2</sub>O and adipic acid by salting out with K<sub>2</sub>CO<sub>3</sub>, washed with H<sub>2</sub>O, dried over CaCl<sub>2</sub> or anhydrous K<sub>2</sub>CO<sub>3</sub>, filtered and fractionated through an efficient column; fraction with b 128-131°/atm of the pure *ketone* (86-92g, 75-80%) is collected. [Thorpe & Kon *Org Synth Coll Vol* **1** 192 1932.] If cyclopentanone has been standing for a while, it can be purified by shaking it with aqueous KMnO<sub>4</sub> to remove materials absorbing around 230 to 240nm. Dry it over Linde-type 13X molecular sieves and fractionally distil it. It has also been purified by conversion to the NaHSO<sub>3</sub> adduct which, after crystallising four times from EtOH/water (4:1), is decomposed by adding to an equal weight of Na<sub>2</sub>CO<sub>3</sub> in hot H<sub>2</sub>O. The free cyclopentanone is steam distilled from the solution. The distillate is saturated with NaCl and extracted with \*benzene (do not use H<sub>2</sub>O, see above) which is then dried (anhydrous K<sub>2</sub>CO<sub>3</sub>), filtered and evaporated. The residue is then distilled [Allen, et al. *J Chem Soc* 1909 1960]. Its FT-IR (NaCl) has  $\nu_{\max}$  at 2966 (CH), 1746.4 (C=O), 1407.6 (OH enol ?), 1278.2, 1153.0, 959.2, 834.2,

582.2 and 471.7  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (15MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.97 (t, 4H, C-3 and C-4 -methylenes), 2.17 (t, 4H, C-2 and C-5 -methylenes); and its  $^{13}\text{C}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 23.24 (C-3,4), 38.30 (C-2,5) and 220.16 (C-1).

The *oxime* [1192-28-5] forms prisms with **m 56.5°**, **b 120-121°/45mm**, **196°/atm**, the *semicarbazone* [5459-00-7] has **m 224°**, the *2,4-dinitrophenylhydrazone* [2057-87-6] has **m 145.5, 146.5°**, and the *ethylene ketal (1,4-dioxaspiro[4.4]nonane* [176-32-9] has **m 153°**. The *enol*-form is present to the extent of  $2.5 \times 10^{-7}\%$  in the gas phase.

**Cyclopentyl methyl ether (CPME)** [5614-37-9] **M 100.2**, **fp -134.8°**, **b 105.44°/760mm 105-106.5°/760mm 106°/760mm**,  $d_4^{20}$  **0.8627**,  $n_D^{20}$  **1.4206**. This ether has the advantage of having a high resistance to hydroperoxide formation. It is an “environmentally friendly” solvent and is an alternative to ether solvents like THF,  $\text{Et}_2\text{O}$  and methyl *tert*-butyl ether. Two methods after Williamson’s synthesis have been reported. In the *first* Na (370g, 16g. atoms) is dissolved in MeOH (2370g, 74moles) and excess of MeOH is distilled off until NaOMe starts to separate. The solution is cooled to 60° and cyclopentyl chloride (1254g, 12moles, b 114°/atm [930-28-9]) is added slowly and kept at 60° for 100 hours. The mixture is then fractionated to give recovered cyclopentyl chloride (21%), cyclopentyl methyl ether (278g, 29% based on reacted cyclopentyl chloride) and cyclopentene (31% which distils as an azeotrope with MeOH (b 37-38°/atm). [Olson et al. *J Am Chem Soc* **69** 2451 1947.] In the *second*, Na (15g) is dispersed under hot xylene, cooled, the xylene is replaced with anhydrous  $\text{Et}_2\text{O}$  (150ml), and a solution of cyclopentanol (57g, b 141-142°/769mm, see [96-41-3]) in dry  $\text{Et}_2\text{O}$  (~100ml) is added with stirring during 3 hours and allowed to stand for 12 hours. Methyl iodide (103g) is then added during 2 hours to the preceding sodio compound as the  $\text{Et}_2\text{O}$  boils gently, allowed to stand overnight, the  $\text{Et}_2\text{O}$  is distilled off, the crude CPME is then distilled out and redistilled under  $\text{N}_2$  at 105°/763mm to give pure *cyclopentyl methyl ether* (21g, 32%). It can be stabilised with 50ppm of BHT. [Vogel *J Chem Soc* 1809 1948, *Beilstein* **6** III 5, **6** IV 6.]

**Cyclotetradecane** [295-17-0] **M 192.3**, **m 56°**. Recrystallise it twice from aqueous EtOH, then sublime it *in vacuo* [Dretloff et al. *J Am Chem Soc* **109** 7797 1987].

**Cyclotetradecanone** [3603-99-4] **M 206.3**, **m 25°**, **b 145°/10mm**,  $d_4^{20}$  **0.926**,  $n_D^{20}$  **1.480**. It is converted to the semicarbazone which is recrystallised from EtOH and re-converted to the free cyclotetradecanone by hydrolysis [Dretloff et al. *J Am Chem Soc* **109** 7797 1987]. Fractionate it in a vacuum.

**Decahydronaphthalene (decalin, mixed isomers)** [91-17-8] **M 138.2**, **b 191.7°/760mm**,  $d_4^{20}$  **0.886**,  $n_D^{20}$  **1.476**. Stir decalin with conc  $\text{H}_2\text{SO}_4$  for several hours. Then the organic phase is separated, washed with water, saturated aqueous  $\text{Na}_2\text{CO}_3$ , again with water, dried with  $\text{CaSO}_4$  or  $\text{CaH}_2$  (and perhaps dried further with Na), filtered and distilled under reduced pressure (**b 63-70°/10mm**). It has also been purified by repeated passage through long columns of silica gel previously activated at 200-250°, followed by distillation from  $\text{LiAlH}_4$  and storage under  $\text{N}_2$ . Type 4A molecular sieves can be used as drying agent. Storage over silica gel removes water and other polar substances. [For the separation of *cis* and *trans* isomers see Seyer & Walker *J Am Chem Soc* **60** 2125 1938, and Baker & Schuetz *J Am Chem Soc* **69** 1250 1949.]

***cis*-Decahydronaphthalene** [493-01-6] **M 138.2**, **f -43.2°**, **b 195.7°, 81-83°/19mm**,  $d_4^{20}$  **0.897**,  $n_D^{20}$  **1.48113**. Purification methods described for the mixed isomers are applicable here. The individual isomers can be separated by very efficient fractional distillation, followed by fractional crystallisation by partial freezing. The *cis*-isomer reacts preferentially with  $\text{AlCl}_3$  and can be removed from the *trans*-isomer by stirring the mixture with a limited amount of  $\text{AlCl}_3$  for 48 hours at room temperature, filtering and distilling. [Seyer & Walker *J Am Chem Soc* **60** 2125 1938, Baker & Schuetz *J Am Chem Soc* **69** 1250 1949.] A very pure authentic sample is best obtained by synthesis from *cis*-1,2-bis-chloroethylcyclohexane [Whitesides & Gutowski *J Org Chem* **41** 2882 1976, *Beilstein* **5** IV 310.]

***trans*-Decahydronaphthalene** [493-02-7] **M 138.2**, **f -30.6°**, **b 187.3°**,  $d_4^{20}$  **0.870**,  $n_D^{20}$  **1.46968**. See purification of *cis*-isomer above. [Seyer & Walker *J Am Chem Soc* **60** 2125 1938, Baker & Schuetz *J Am Chem Soc* **69** 1250 1949, *Beilstein* **5** IV 311.]

(+)-Dehydroabietylamine (abieta-8,11,13-triene-18-ylamine) [1446-61-3] **M 285.5, m 41°, 42.5-45°, b 192-193°/1mm, 250°/12mm, n<sub>D</sub><sup>40</sup> 1.546, [α]<sub>546</sub><sup>20</sup> +51° (c 1, EtOH), pK<sub>Est</sub> ~10.3.** The crude base is purified by converting 2g of base in toluene (3.3ml) into the *acetate salt* by heating at 65-70° with 0.46g of AcOH, and the crystals are collected and dried (0.96g from two crops, **m 141-143°**). The acetate salt is then dissolved in warm H<sub>2</sub>O, basified with aqueous NaOH and extracted with \*C<sub>6</sub>H<sub>6</sub>. The dried extract (MgSO<sub>4</sub>) is evaporated in vacuum leaving a viscous oil which crystallises and can be distilled. [Gottstein & Cheney *J Org Chem* **30** 2072 1965.] The *picrate* has **m 234-236°** (from aqueous MeOH), and the *formate* has **m 147-148°** (from heptane). [Beilstein **12** IV 3005.]

**Diamantane (congressane)** [2292-79-7] **M 188.3, m 234-235°, 243-245°.** Purify diamantane by repeated crystallisation from MeOH or pentane. *Alternatively*, purify it by dissolving it in CH<sub>2</sub>Cl<sub>2</sub>, washing with 5% aqueous NaOH and water, and drying (MgSO<sub>4</sub>). The solution is filtered, concentrated to a small volume, an equal weight of alumina is added, and the solvent evaporated. The residue is placed on an activated alumina column (ca 4 x weight of diamantane) and eluted with petroleum ether (b 40-60°). Eight sublimations and twenty zone refining experiments gave material **m 251°** of 99.99% purity by differential analysis [Gund et al. *Tetrahedron Lett* 3877 1970, Courtney et al. *J Chem Soc Perkin Trans I* 2691 1972]. [For spectra see Cupas et al. *J Am Chem Soc* **87** 919 1965.]

**1,3-Diaminoadamantane** [10303-95-4] **M 164.3, m 52°, pK<sub>Est(1)</sub> ~8.6, pK<sub>Est(2)</sub> ~10.6.** Purify it by zone refining. The *dibenzoyl* derivative has **m 248°**, and the *dihydrochloride salt* has **m 310° (360°)** after recrystallisation from aqueous EtOH or EtOH/Et<sub>2</sub>O. [Prelog & Seiwert *Chem Ber* **74** 1769 1941, Stetter & Wulff *Chem Ber* **93** 1366 1960.]

**cis-1,2-Diaminocyclohexane (meso-chxn)** [1436-59-5] **M 114.2, b 92-93°/8mm, 170°/atm, d<sub>4</sub><sup>25</sup> 0.952, n<sub>D</sub><sup>20</sup> 1.493, pK<sub>1</sub><sup>20</sup> 6.13 (6.41), pK<sub>2</sub><sup>20</sup> 9.93(9.91).** It is prepared by reduction of the *cis*-1,2-dioxime with boiling Na/EtOH, or from *cis*-1,2-dihydrazinocarbonylcyclohexane with NaNO<sub>2</sub>/HCl [Jaeger & Blumendal *Z Anorg Chem* **175** 165 1928]. Dry the diamine over solid KOH and distil it in a vacuum. It is a strong base, keep it away from CO<sub>2</sub>, and store it in the dark under N<sub>2</sub>. The *dihydrochloride* has **m 307-310°**, and the *dipicrate* (from H<sub>2</sub>O) has **m 260°(dec)**. This *cis*-diamine is a *meso* form because its mirror image is identical to it, and is superimposable over it (see below) [Beilstein **13** II 3, **13** IV 5.]

**(±)-trans-1,2-Diaminocyclohexane [(±)-chxn]** [1121-22-8] **M 114.2, m 14-15°, b 78-81°/15mm, 85-88°/25mm, d<sub>4</sub><sup>25</sup> 0.951, n<sub>D</sub><sup>20</sup> 1.489, pK<sub>1</sub><sup>20</sup> 6.47(6.72), pK<sub>2</sub><sup>20</sup> 9.94(9.86).** Purify this *racemic base*, and store it as for the *cis*-isomer above since it is a strong base, and becomes yellow on storage. [Beilstein **13** H 1, **13** III 8, **13** IV 5.]

**cis/trans-1,2-Diaminocyclohexane resolution.** [Note that the *trans*- isomer is a *racemic mixture*, whereas the *cis*- isomer has a *meso configuration* and its mirror image is identical and superimposable over it.] A solution of L(+)-tartaric acid (150g, 0.99mol) in distilled H<sub>2</sub>O (400ml) in a 1L beaker is stirred until clear, and a mixture of *cis*- and *trans*- 1,2-diaminocyclohexane (240ml, 1.94mol, of ~60/40 *trans/cis* mixture which is commercially available and is cheaper than the pure *trans* compound for which the same procedure applies) is added at a rate whereby the temperature rises to 70°. Glacial acetic acid (100ml, 1.75mol) is then added at such a rate that the temperature rises to 90°. A white precipitate immediately separates as the acid is added, and the slurry is stirred vigorously until the temperature cools slowly to 25° (~2 hours), and then is cooled further to ~5° in an ice bath for 2 hours, and is collected by filtration. The filter cake is washed with H<sub>2</sub>O (100ml at ~5°), then EtOH (5 x 100ml), and dried by sucking air through it for 1 hour and analysed for enantiomeric purity *via* the bis-*m*-toluoyl amide\*. The salt is then dried *in vacuo* at ~40° to give *R,R*-1,2-diamoniumcyclohexane mono-(+)-tartrate as a white solid (160g, 99%) in ≥99% enantiomeric excess (ee). If the optical purity is not as expected then recrystallise the salt in two crops from Me<sub>2</sub>CO (1:20 w/v with ~60-70% recovery). The *S,S*-diastereomeric tartrate salt is obtained by using D(-)-tartaric acid. If the 'ee' is <99%, or if the difference between the top and the bottom of the cake is >0.2%, then the cake should be washed with more MeOH. If product of ≥99% 'ee' is required then recrystallisation of the tartrate salt by dissolving it in H<sub>2</sub>O (≈ 1:10 w/v) with heating to 90°, followed by cooling to 5° overnight, gives ~60-70% recovery.

\* The diammonium salt (25mg) is mixed with 4N NaOH (0.5ml), CH<sub>2</sub>Cl<sub>2</sub> (1.5ml) and *m*-toluoyl chloride (50μl)

with vigorous mixing. The lower (organic layer) is diluted with 10 $\mu$ l of *iso*-propanol, and 10 $\mu$ l of this is analysed by HPLC using *iso*-propanol/hexane (1:9) mixture at 1ml/minute flow rate. [Larrow et al. *J Org Chem* **59** 1939 1994, Gasbøl et al. *Acad Chem Scand* **26** 3605 1972.]

**1*R*,2*R*(-)-trans-1,2-Diaminocyclohexane** [(-)-chxn] [20439-47-8] **M 114.2, m 41-45 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -25.5 $^{\circ}$  (c 5 M HCl), m 14-15 $^{\circ}$** . Distil or recrystallise the diamine from petroleum ether under N<sub>2</sub> or Ar. It has a plain-negative ORD curve [Gillard *Tetrahedron* **21** 503 1965, O'Brien & Toole *J Am Chem Soc* **77** 1368 1955]. Store it as above. The *1*R*,2*R*-base L-tartrate salt* has [39961-95-0], **M 264.3, m 273 $^{\circ}$  and  $[\alpha]_{\text{D}}^{25}$  +12.5 $^{\circ}$  (c 4, H<sub>2</sub>O)**, and can be used to purify and/or optically enrich the free base. [Beilstein **13** III 6, and references below.]

**1*S*,2*S*(+)-trans-1,2-Diaminocyclohexane** [(+)-chxn] [21436-03-3] **M 114.2, m 42-45 $^{\circ}$ , b 104-110 $^{\circ}$ /40mm,  $[\alpha]_{\text{D}}^{20}$  +25.5 $^{\circ}$  (c 5 M HCl)**. Distil or recrystallise the diamine from petroleum ether under N<sub>2</sub> or Ar. Store it as above. It has a plain-positive ORD curve [Gillard *Tetrahedron* **21** 503 1965, O'Brien & Toole *J Am Chem Soc* **77** 1368 1955]. The *1*S*,2*S*-base D-tartrate salt* has [67333-70-4], **M 264.3, m 180-184 $^{\circ}$ (dec) and  $[\alpha]_{\text{D}}^{25}$  -12.5 $^{\circ}$  (c 4, H<sub>2</sub>O)** from which the free base can be purified or optically enriched. It is a useful chiral synthon. [Fjii et al. *J Chem Soc, Chem Commun* **45** 1985, Takahashi et al. *Tetrahedron Lett* **30** 7095 1989, Hanassian et al. *J Org Chem* **58** 1991 1993, for absolute configuration see Gillard *Tetrahedron* **21** 503 1965, Beilstein **13** III 7.]

**trans-1,4-Diaminocyclohexane** [2615-25-0] **M 114.2, m 69-72 $^{\circ}$ , b 197 $^{\circ}$ /760mm, pK<sub>est(1)</sub> ~9.4, pK<sub>est(2)</sub> 10.8**. Recrystallise the diamine from petroleum ether under N<sub>2</sub> or Ar as it should be an even stronger base than the above 1,2-diamine isomers. It distils under N<sub>2</sub>. Store it in the dark under N<sub>2</sub>. [Beilstein **13** I 3, **13** III 11.]

**(1*R*,2*S*)-cis-1,2-Diaminocyclopentane** (**1*R*,2*S**R*-cis-1,2-cyclopentanediamine, meso-1,2-cyclopentanediamine, (meso)-cptn**) [40535-45-3; no configuration 41330-23-8] **M 100.1, b 62-65 $^{\circ}$ /13.5mm, 99 $^{\circ}$ /60mm**. Whereas reduction of *anti*-cyclopentane-1,2-dione dioxime (see below, and from reaction of cyclopentane-1,2-dione and NH<sub>2</sub>OH) with Na/EtOH yields predominantly *trans*-( $\pm$ )-diamine, the reaction of *cyclopentane-1,2-dione monooxime* [31597-37-0] with NH<sub>2</sub>OH gives a 2:1 mixture of *anti*- and *amphi*-dioximes which provide the less water soluble red bis-*anti*-dioximateNi(II) complex and the more water soluble yellow-brown bis-*amphi*-dioximateNi(II) complex respectively. When this mixture of Ni(II) complexes, or the *amphi*-complex, is reduced with KBH<sub>4</sub> in "diglyme", the base finally isolated is pure *cis*-1,2-cyclopentanediamine.

Thus, *bis(amphi-cyclopentane-1,2-dione dioximate)nickel(II)* (79g, 90.5mol, dried over H<sub>2</sub>SO<sub>4</sub> *in vacuo*) is added to a solution of KBH<sub>4</sub> (30g) dissolved in dry diglyme [600ml, bis(2-methoxyethyl) ether, [111-96-6]] in 2L flask. N<sub>2</sub> is bubbled through the solution, then during 1 hour anhydrous AlCl<sub>3</sub> (25g) in dry diglyme (100ml) is added dropwise carefully as vigorous evolution of H<sub>2</sub> ensues; the temperature being kept below 35 $^{\circ}$  by cooling in ice-water. When evolution of H<sub>2</sub> ceases, the mixture is heated at 70 $^{\circ}$  for 18 hours. Then a solution of KOH (100g) in H<sub>2</sub>O (150ml) is added carefully, and the *cis*-diamine is steam distilled off; 6L of distillate are collected, acidified with 12M HCl (to pH 3), and the dry *cis*-1,2-cyclopentanediamine dihydrochloride (23g, 27% based on the Ni complex) is isolated as for the *trans*-( $\pm$ )-isomer below. The free base, *cis*-1,2-cyclopentanediamine **b 62-65 $^{\circ}$ /13.5mm** (83%, based on the dihydrochloride) is stored under N<sub>2</sub> at -20 $^{\circ}$ . [Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972]. This *cis*-diamine has been fractionated through a vacuum-jacketed distillation column as described by Ray et al. *Rev Sci Instr* **28** 200 1957, and the fraction with **b 99 $^{\circ}$ /60mm** contained pure *cis*-isomer [Phillips & Royer *Inorg Chem* **4** 616 1965].

The <sup>1</sup>H NMR in D<sub>2</sub>O (TMS as external standard) can be used to distinguish between *trans*-( $\pm$ )-1,2-diamine and the preceding *cis*-(*meso*)-1,2-diamine because the triplets from carbon atoms 1 + 2 occur at  $\delta$  2.72 and  $\delta$  2.77 respectively [Phillips & Royer *Inorg Chem* **4** 616 1965].

Distinct bis- and tris- *cis*-cptn complexes such as [Ni(*cis*-cptn)<sub>2</sub>](ClO<sub>4</sub>)<sub>2</sub>·6H<sub>2</sub>O, [Ni(*cis*-cptn)<sub>3</sub>](ClO<sub>4</sub>)<sub>2</sub> and [Co(*cis*-cptn)<sub>3</sub>]Br<sub>2</sub>·3H<sub>2</sub>O are readily formed [Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972]. [Beilstein **4** II 3.]

**(1*R*,2*R*)-( $\pm$ )-trans-1,2-Diaminocyclopentane** [(**1*R*,2*R*)-trans-1,2-cyclopentanediamine, ( $\pm$ )-trans-1,2-cyclopentanediamine, ( $\pm$ )-cptn] [3145-88-8] **M 100.1, b 65-67 $^{\circ}$ /13.5mm, 103 $^{\circ}$ /60mm, 170 $^{\circ}$ /atm, pK<sub>Est 1</sub> ~7.0, pK<sub>Est 2</sub> ~10.0**. This diamine is a strong base and is best kept as its salts which are not greatly affected by moisture and CO<sub>2</sub>. It is prepared by reduction of *anti*-cyclopentane-1,2-dione dioxime [6635-29-6] **M 128.1, m 210 $^{\circ}$  (215 $^{\circ}$  dec; 210-225 $^{\circ}$  dec, darkening at 192-205 $^{\circ}$  also reported for an analytically pure sample) which is****

obtained by adding *2-ethoxycarbonylcyclopentanone* (31.2g, b 102-104°/12mm, [611-10-9]) to a solution of NaOH (8.8g) in H<sub>2</sub>O (160ml) whereby a heavy white precipitate separates. A solution of NaNO<sub>2</sub> (13.8g) in H<sub>2</sub>O (32ml) is then added to it and the mixture, under N<sub>2</sub>, is shaken mechanically for 44 hours at ~25°. The resulting clear yellow solution is cooled to 0°, and at this temperature 6N H<sub>2</sub>SO<sub>4</sub> is added carefully while CO<sub>2</sub>, and oxides of nitrogen evolve. The solution is extracted with Et<sub>2</sub>O (in a continuous liquid-liquid extractor) for 5.5 hours, the extract is evaporated under N<sub>2</sub> and reduced pressure to a small volume of ~20ml. This suspension of crude cyclopentane-1,2-dione monooximes is treated with a solution of hydroxylamine hydrochloride (13.9g) and NaOH (8.0g) in H<sub>2</sub>O (50ml) when heat is evolved and a precipitate separates immediately. The mixture is allowed to stand overnight at 5°, the light tan crude *anti-cyclopentane-1,2-dione dioxime* (16.8g, 66%, decomposes at 230-240°) is filtered off and is dried *in vacuo*, leaving the more soluble *amphi-cyclopentane-1,2-dione dioxime* in the filtrate. The analytically pure *anti-dioxime* (dec at 210-225° without melting, 83% recovery, [6635-29-6]) can be obtained by dissolving the crude material (1g) in 2% aqueous NaOH (25ml) at ~25° (slightly turbid), stirring with Norit, filtering, neutralising the filtrate with 1 N HCl (to phenolphthalein), filtering off the solid, washing it twice with H<sub>2</sub>O, once with Me<sub>2</sub>CO and it drying *in vacuo*. [Cope et al. *J Am Chem Soc* **73** 1199 1951, Lloyd & Marshall *J Chem Soc* 2597 1956.]

*Alternatively*, to a stirred mixture of granulated Na (38g) and \*C<sub>6</sub>H<sub>6</sub> (900ml, dried over Na) is added diethyl adipate (225ml, 1.11 mole, [141-28-6]) rapidly followed by absolute EtOH (5ml), and is boiled under reflux with stirring for 18 hours then cooled to 0°. The white solid that separated is filtered off, washed with a little \*C<sub>6</sub>H<sub>6</sub> and dried *in vacuo* over shredded paraffin wax to give *sodium cyclopentanone-2-carboxylate* (180g). This salt (1mol) is partly dissolved in ice-water (600ml), N<sub>2</sub> is bubbled through for 15 minutes to remove O<sub>2</sub>; and a mixture of a solution of NaOH (4g, 0.1mol) and NaNO<sub>2</sub> (70g, 1.03mol) in H<sub>2</sub>O (160ml) is added dropwise to it. The mixture is then heated under N<sub>2</sub> with vigorous stirring at 40° for 30 hours. After cooling to 0°, 6M H<sub>2</sub>SO<sub>4</sub> (200ml, 1.2mol) is added in one hour; and when evolution of CO<sub>2</sub> has ceased the solution is neutralised with 12M ammonia. A solution of NH<sub>2</sub>OH HCl (70g, 1.0mol) in H<sub>2</sub>O (160ml), neutralised with K<sub>2</sub>CO<sub>3</sub> (70g, 0.5mol), is immediately added. After 1 hour, the precipitation of crude *anti-cyclopentane-1,2-dione dioxime* as light tanned needle-like crystals is complete; it is filtered off, washed with a little H<sub>2</sub>O and dried *in vacuo* over H<sub>2</sub>SO<sub>4</sub> for 48 hours (yield 90g, 62% calculated on diethyl adipate used). The filtrate contains the *amphi-cyclopentane-1,2-dione dioxime* which is isolated as the yellow-brown *bis(amphi-cyclopentane-1,2-dione dioximato)nickel(II) complex* (30g) by slowly adding Ni(II)SO<sub>4</sub> 7H<sub>2</sub>O (20g) in H<sub>2</sub>O (60ml) to this filtrate, and washing the complex with H<sub>2</sub>O and drying in air. [Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972.]

*Bis(amphi-cyclopentane-1,2-dione dioximato)nickel(II) complex* is obtained analytically pure, as a red precipitate in 70% yield, from *anti-cyclopentane-1,2-dione dioxime* (12.8g, 0.1mol) in 4M ammonia (15ml), by adding slowly Ni(II)Cl<sub>2</sub> 6H<sub>2</sub>O (12g, 0.05mol) in H<sub>2</sub>O (40ml) to it, filtering off the solid, washing it with H<sub>2</sub>O and drying it in air. [Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972.]

*Reduction to (±)-diamine*: The *anti-cyclopentane-1,2-dione dioxime* (10g) in absolute EtOH (2L) is reduced by addition of freshly cut small pieces of Na metal (180g) over a period of several hours. The solution is refluxed until all the metal has dissolved. The EtOH is distilled off *in vacuo*, the residue is basified strongly with 10N NaOH, and the base is steam distilled. The distillate is made strongly alkaline with aqueous NaOH, extracted with trichloroethylene, the solvent is evaporated and *trans-1,2-cyclopentanediamine* then distils at **65°/13.5mm**,  $n_D^{25}$  **1.4850-1.4858** in 29% yield. When the fractional distillation is carried out using a column (90 x 1 cm) packed with glass helices and the reflux ratio is kept at >40:1, the first 25% of distillate contains a significant amount of lower boiling *cis-diamine*, **99°/60mm**; the remaining 75% of material then boils at a constant temperature, **103°/60mm**, and is the *trans-(±)-diamine*. *Alternatively*, the steam distillate is acidified with concentrated HCl, the solution is evaporated to dryness *in vacuo*, and the residue is recrystallised from EtOH/Et<sub>2</sub>O to provide *trans-(±)-1,2-cyclopentanediamine dihydrochloride* in 77% yield, which **decomposes at 287-290°** (darkening gradually >230°), and forms a *dihydrate* on exposure to air. The *trans-(±)-dipicrate* has **m 233-233.5° (dec)** after recrystallisation from aqueous EtOH or H<sub>2</sub>O, and the *trans-(±)-diacetyl derivative* has **m 226.5-227.5° (dec)** after recrystallisation from EtOH by addition of Et<sub>2</sub>O. [Cope et al. *J Am Chem Soc* **73** 1199 1951, Phillips & Royer *Inorg Chem* **4** 616 1965, Lloyd & Marshall *J Chem Soc* 2599 1956, Jaeger & Blumendal *Z Anorg Allgem Chem* **175** 161 1928.]

*Alternative reduction to (±)-diamine*: The *anti-cyclopentane-1,2-dione dioxime* (20g, 156mmol) in absolute EtOH (1.5L) is reduced by heating to 40° (internal temperature), and stirring under a reflux condenser, with Mg powder (8g) and Hg(II)Cl<sub>2</sub> (0.1g) while Na metal (150g, 6.5mol, in 5g pieces) is added *via* the top of the condenser and keeping the temperature below 50° (internal temperature). The mixture is finally heated to boil-

ing in order to dissolve all the Na. The ( $\pm$ )-*diamine* is isolated by steam distillation (external heating being necessary so as to maintain a small volume in the flask) until the pH of the distillate is below 8 (*ca* 4L). This is acidified (pH 3) with 12M HCl, evaporated almost to dryness *in vacuo* (at  $\sim 25^\circ$ ), excess HCl is removed by washing with EtOH/Et<sub>2</sub>O (2:1v/v, 20ml), and the hygroscopic salt is dried over KOH in a desiccator to give the ( $\pm$ )-*diamine dihydrochloride* (5g, 18% based on dioxime). The free base is obtained by adding the dihydrochloride (5g, 29mmol) in small portions with stirring to a 1:1v/v mixture of MeOH/Et<sub>2</sub>O (50ml) into which is dissolved Na (1.7g, 74mmol, cooling). After stirring for 1 hour, the precipitated NaCl is filtered off, extracted with MeOH/Et<sub>2</sub>O (1:1v/v, 5ml), the combined extract and filtrate are evaporated to  $\sim 3$ ml and this residue is distilled in a N<sub>2</sub> atmosphere to give ( $\pm$ )-*trans-cyclopentane-1,2-diamine*, **b 65-67<sup>o</sup>/13.5mm** (2.5g, 89%, based on the dihydrochloride) which is stored under N<sub>2</sub> at  $-20^\circ$ . Distinct *trans*-( $\pm$ )-cptn complexes of Ni, Co, Rh and Pt have been prepared; some are enantiomeric and others are structural isomers [Phillips & Royer *Inorg Chem* **4** 616 1965, Jaeger & Blumendal *Z Anorg Allgem Chem* **175** 161 1928, Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972].

The <sup>1</sup>H NMR in D<sub>2</sub>O (TMS as external standard) can be used to distinguish between *trans*-( $\pm$ )-1,2-diamine and the preceding *cis*-(*meso*)-1,2-diamine because the triplets from carbon atoms 1 + 2 occur at  $\delta$  2.72 and  $\delta$  2.77 respectively [Phillips & Royer *Inorg Chem* **4** 616 1965].

**(1R,2R)-(-)-trans-1,2-Diaminocyclopentane dihydrochloride [(-)-cptn HCl]** [1030390-38-5] **M 173.1, m dec**,  $[\alpha]_D^{25} -23.5^\circ$  (**c 2, 1M HCl**) and **(1S,2S)-(+)-trans-1,2-Diaminocyclopentane dihydrochloride [(+)-cptn HCl]** [477873-22-6] **M 173.1, m dec**,  $[\alpha]_D^{25} +23^\circ$  (**c 2, 1M HCl**). **Optical resolution:** ( $\pm$ )-Cptn (2g, 20mmol) is resolved by adding it to a solution of D-(+)-tartaric acid (7g, 47mmol) in MeOH (1ml) and H<sub>2</sub>O (7ml), and at 40<sup>o</sup> MeOH (12ml) is added dropwise. The mixture is then kept at 3<sup>o</sup> for 18 hours, the crystalline solid is filtered off, washed with MeOH (10ml) to give the less soluble hydrogen tartrate salt [5.2g, 65% based on ( $\pm$ )-cptn]. After two recrystallisations, the **(-)-cptn-(+)-tartrate salt** [2.5g, 31.4% based on ( $\pm$ )-cptn] has **m 134-135<sup>o</sup>** (143-144<sup>o</sup> also reported),  $[\alpha]_D^{25} +10.1^\circ$  (**c 2, H<sub>2</sub>O**). When this salt (2.4g, 6mmol) in H<sub>2</sub>O (8ml) is treated with a solution of KCl (0.9g, 12mmol) in H<sub>2</sub>O (4ml), the highly insoluble potassium hydrogen tartrate separates and is filtered off after standing for 1 hour. Addition of solid KOH (8g, with cooling) allowed solid (-)-cptn hydrate to separate, but the mixture is extracted thoroughly with Et<sub>2</sub>O (5 x 10ml), the extract is dried (KOH or MgSO<sub>4</sub>), filtered, and evaporated to give **(-)-trans-cptn** (0.55g, 92% based on tartrate salt). When carried out on a larger scale the **(-)-base** can be distilled, **b 166<sup>o</sup>/atm, d<sup>25</sup> 0.9463**, and has  $[\alpha]_D -54.7^\circ$  (**neat**),  $[\alpha]_D -42^\circ$  (**\*C<sub>6</sub>H<sub>6</sub>**),  $[\alpha]_D -39^\circ$  (**H<sub>2</sub>O**), and a plain-negative ORD curve. **(-)-trans-Cptn sulfate** has  $[\alpha]_D -12^\circ$  (**H<sub>2</sub>O**). [Toftlund & Pedersen *Acta Chim Scand* **26** 4019 1972, Jaeger & Blumendal *Z Anorg Allgem Chem* **175** 161 1928, Phillips & Royer *Inorg Chem* **4** 616 1965]. **(+)-Cptn-(+)-tartrate** can be obtained from the mother liquors of the above resolution in order to isolate the (+)-base; but it is advisable to collect all the mother liquors, liberate the free base with KOH, isolate it and repeat the process using L-(-)-tartaric acid to form the less soluble diastereoisomeric **(+)-cptn-(-)-tartrate salt**. Pure **(+)-trans-cptn** can then be isolated as above and will have the same physical properties except for the optical rotations that will be of opposite sign and has a plain-positive ORD curve [Dunlop et al. *J Chem Soc* 3160 1964, Gillard *Tetrahedron* **21** 503 1965, O'Brien & Toole *J Am Chem Soc* **77** 1368 1955]. [Beilstein **4** II 3, **4** III 5]

Complexes of optically active cptn with Rh, Co, Ni and Pt have been prepared, and in some cases enantiomeric and geometrical isomers have been identified. One has been used to obtain pure enantiomeric diamine, for example, the complex **[Co<sub>2</sub>( $\pm$ -cptn)<sub>7</sub> (H<sub>2</sub>O)<sub>2</sub>]<sup>6+</sup>** has been resolved into three bands on Amberlite CG-50 carboxylic acid resin 400-400 by elution with aqueous HCl (pH 2, flow rate 5ml/hour) and collecting 40ml fractions. Bands I, II and III are probably geometrical isomers because when the isolated complexes are decomposed with concentrated aqueous NaOH, extracted with \*C<sub>6</sub>H<sub>6</sub>, the extract dried and evaporated, the free base **(+)-cptn** with  $[\alpha]_D -42^\circ \pm 1^\circ$  (**\*C<sub>6</sub>H<sub>6</sub>**) is obtained (compare with above) [Phillips & Royer *Inorg Chem* **4** 616 1965].

The optical properties (ORD, CD) of several cptn-transition metal complexes have been thoroughly investigated and used to determine and confirm the absolute configuration of cptn [Dunlop et al. *J Chem Soc* 3160 1964, Gillard *Tetrahedron* **21** 503 1965]. The crystal structure of **(-)<sub>587</sub>-[Co(+trans-cptn)<sub>3</sub>]Cl<sub>3</sub>·4H<sub>2</sub>O** confirms the **1S,2S-configuration** for the dextro-enantiomer [Ito et al. *Acta Cryst* **B** **27** 2187 1971].

**trans-1,2-Dibromocyclopentane** [10230-26-9] **M 227.9, b 72.5<sup>o</sup>/15mm, 74<sup>o</sup>/17mm, d<sub>4</sub><sup>20</sup> 1.857, n<sub>D</sub><sup>25</sup> 1.5460**. It is prepared by addition of bromine to cyclopentene at  $-20^\circ$  followed by fractional distillation. The <sup>1</sup>H NMR

(60MHz, CCl<sub>4</sub>) has a complex multiplet centered at ~2.5 ppm (3 x CH<sub>2</sub>), and two triplets centered at 4.58ppm ( $J = 7.4\text{Hz}$ , for 2 x CHBr) from TMS. [Altona et al. *Rec Trav Chim, Pays Bas* **85** 983 1966, cf Abell et al. *J Am Chem Soc* **82** 3610 1960.]

**cis-3,4-Dichlorocyclobutene** [2957-95-1] **M 123.0, b 70-71°/55mm, 74-76°/55mm, d<sub>4</sub><sup>20</sup> 1.297, n<sub>D</sub><sup>20</sup> 1.499.** Distil the cyclobutene at 55mm through a 36-in platinum spinning band column, a fore-run b 58-62°/55mm is mainly 1,4-dichlorobutadiene. When the temperature reaches 70° the reflux ratio is reduced to 10:1 and the product is collected quickly. It is usually necessary to apply heat frequently with a sun lamp to prevent any dichlorobutadiene from clogging the exit in the early part of the distillation [Pettit & Henery *Org Synth* **50** 36 1970].

**Dicyclohexylamine (Cy<sub>2</sub>NH<sub>2</sub>)** [101-83-7] **M 181.3, m 20°, b 83°/1mm, 99.3°/4mm, 113.5°/9mm, 117-120°/10mm, 135.4°/30mm, 154.3°/50mm, 199°/200mm, 255.8°/atm, d<sub>4</sub><sup>20</sup> 0.912, n<sub>D</sub><sup>20</sup> 1.4845, pK<sup>25</sup> 11.25.** It is a strong base, has a fishy odour, is soluble in H<sub>2</sub>O, and organic solvents and forms adducts with some of them. It can be purified by fractional distillation at atmospheric pressure, but if the distillate is coloured (e.g. green) then distillation under reduced pressure gives a clear colourless liquid [Vogel *J Chem Soc* 1828 1928]. The *hydrochloride* [4693-92-9] crystallises from H<sub>2</sub>O or EtOH and has **m 334-335°** (sealed tube), and the *acetyl derivative* [1563-91-3] has **m 103°** (from Et<sub>2</sub>O). [Diwoy & Adkins *J Am Chem Soc* **53** 1870 1931, *Beilstein* **12** H 6, **12** I 114, **12** II 7, **12** III 19, **12** IV 22.] **SKIN and EYE IRRITANT.**

**1,3-Dicyclohexylcarbodiimide (DCC)** [538-75-0] **M 206.3, m 34-35°, b 95-97°/0.2mm, 120-121°/0.6mm, 155°/11mm.** It is sampled as a liquid after melting in warm H<sub>2</sub>O. It is sensitive to air, and *it is a potent skin irritant*. It can be distilled in a vacuum, and is best stored in a tightly stoppered bottle in a freezer. It dissolves readily in CH<sub>2</sub>Cl<sub>2</sub> and pyridine where the reaction product with H<sub>2</sub>O, after condensation, is dicyclohexyl urea which is insoluble and can be filtered off. *Alternatively*, dissolve it in CH<sub>2</sub>Cl<sub>2</sub>, add powdered anhydrous MgSO<sub>4</sub>, shake for 4 hours, filter, evaporate and distil the residue at 0.6mm pressure and oil bath temperature of 145°. [Bodansky et al. *Biochemical Preparations* **10**, 122 1963, Schmidt & Seefelder *Justus Liebigs Ann Chem* **571** 83 1951, Schmidt et al. *Justus Liebigs Ann Chem* **612** 11 1958, *Beilstein* **12** IV 72.]

**N,N-Dicyclohexylmethylamine (Cy<sub>2</sub>NMe, N-methyldicyclohexylamine)** [7560-83-0] **M 195.3, b 131-133°/13mm, 135-137°/15mm, 150°/50mm, 265°/atm, d<sub>4</sub><sup>25</sup> 0.912, n<sub>D</sub><sup>25</sup> 1.4900, pK 8.2.** When dicyclohexylamine (90.5g, 0.5mol, see 101-83-7) and 55% aqueous formic acid (121.1ml, 2.5mol) are heated at ~100° and 35% aqueous formaldehyde (97.5g, 11.25mol) is added dropwise over an hour, CO<sub>2</sub> evolution occurs. After gas evolution is complete, ~15 hours, concentrated HCl (1mol) is added and the mixture is evaporated *in vacuo*. The free base is liberated from the residue with strong aqueous NaOH and the oil is extracted with Et<sub>2</sub>O, dried (solid KOH), filtered, the filtrate is evaporated *in vacuo* and the residue is fractionally distilled to give Cy<sub>2</sub>NMe in 51% yield. It was also obtained by alkylating Cy<sub>2</sub>NH with dimethylsulfate. [Hünig & Kiessel *J Prakt Chem* **5** 224 1958.] The *hydrochloride* has **m 154-154°** (from EtOAc), and the *methiodide* has **m 228-229°** (from EtOH/Et<sub>2</sub>O). [*Beilstein* **12** I 115, **12** II 8, **12** III 21, **12** IV 23.]

**3,4-Diethoxy-3-cyclobutene-1,2-dione (diethyl squarate)** [5321-87-8] **M 170.2, b 89-91°/0.4mm, 88-92°/0.4mm, d<sub>4</sub><sup>20</sup> 1.162, n<sub>D</sub><sup>25</sup> 1.5000.** Dissolve the ester in Et<sub>2</sub>O, wash it with Na<sub>2</sub>CO<sub>3</sub>, H<sub>2</sub>O and dry (Na<sub>2</sub>SO<sub>4</sub>) it, filter, evaporate and distil it using a Kugelrohr, or purify it by chromatography. Use a Kieselgel column and elute with 20% Et<sub>2</sub>O/petroleum ether (b 40-60°), then with Et<sub>2</sub>O/petroleum ether (1:1), evaporate and distil the residue *in vacuo*. [Dehmlow & Schell *Chem Ber* **113** 1 1980, Perri & Moore *J Am Chem Soc* **112** 1897 1990; IR: Cohen & Cohen *J Am Chem Soc* **88** 1533 1966] **It can cause severe dermatitis.** [Foland et al. *J Am Chem Soc* **111** 975 1989, Perri et al. *Org Synth* **69** 220 1990].

**N,N-Diethylcyclohexylamine** [91-65-6] **M 155.3, b 85-87°/15mm, 193°/760mm, d<sub>25</sub><sup>25</sup> 0.850, n<sub>D</sub><sup>25</sup> 1.4562, pK<sup>25</sup> 10.72.** Dry the amine with BaO and fractionally distil it. It is a strong base, store away from CO<sub>2</sub>. The *picrate* has **m 98-99°** (from aqueous EtOH) and the *methiodide* has **m 224°** (from Me<sub>2</sub>CO/petroleum ether) [Cadogan *J Chem Soc* 1081 1957.] [Bain & Pollard *J Am Chem Soc* **61** 2704 1939, *Beilstein* **12** H 6, **12** III 14, **12** IV 19.]

**Diethyl cyclopropane-1,1-dicarboxylate** [1559-02-0] **M 186.2, b 94-96°/10mm,  $d_4^{25}$  1.055,  $n_D^{20}$  1.433.** If it is free from OH bands in the IR, then fractionally distil the ester and redistil the middle fraction. Otherwise shake it with aqueous  $\text{NaHCO}_3$ , dry it ( $\text{MgSO}_4$ ), filter and distil as before or re-esterify it. [As synthon see Danishefsky *Acc Chem Res* **12** 66 1979, *Beilstein* **9** I 314, **9** II 512, **9** III 3595, **9** IV 2786.]

**Dimedone (5,5-dimethylcyclohexane-1,3-dione)** [126-81-8] **M 140.2, m 148-149°,  $pK^{25}$  5.27.** Crystallise dimedone from acetone (ca 8ml/g), water or aqueous EtOH. Dry it in air. [Schneider & Todd *Org Synth Coll Vol II* 200 1943, *Beilstein* **7** H 559, **7** IV 1999.]

**cis- and trans-1,4-Dimethylcyclohexane** [589-90-2] **M 112.2, b 120°,  $d_4^{20}$  0.788,  $n_D^{25}$  1.427.** Free it from olefins by shaking with conc  $\text{H}_2\text{SO}_4$ , washing with water, drying and fractionally distilling it. [Haggis & Owen *J Chem Soc* 411 1953, *Beilstein* **5** III 102, **5** IV 122.]

**1,2-Dimethylcyclohexene** [1674-10-8] **M 110.2, b 135-136°/760mm,  $d_4^{25}$  0.826,  $n_D^{25}$  1.4587.** Pass it through a column of basic alumina and distil it. If removal of 2-methylmethylenecyclohexane or 2,3-dimethylcyclohexene is required, then fractionation through a centre-rod column operating at ~50 theoretical plates is required. [Hammond & Nevitt *J Am Chem Soc* **76** 4121 1954, *Beilstein* **5** III 213, **5** IV 268.]

**Ethyl chrysanthemumate (ethyl  $\pm$ 2,2-dimethyl-3{c and t}-[2-methylpropenyl]cyclopropane carboxylate)** [97-41-6] **M 196.3, b 98-102°/11mm, 117-121°/20mm.** Purify the ester by vacuum distillation. The free *trans-acid* has **m 54°** (from EtOAc), and the free *cis-acid* has **m 113-116°** (from EtOAc). The *4-nitrophenyl ester* has **m 44-45°** (from petroleum ether) [Campbell & Harper *J Chem Soc* 283 1945, IR: Allen et al. *J Org Chem* **22** 1291 1957]. [*Beilstein* **9** II 45.]

**Ethylcyclohexane** [1678-91-7] **M 112.2, b 131.2°/742mm,  $d_4^{20}$  0.7839,  $n_D^{20}$  1.43304,  $n_D^{25}$  1.43073.** Purify it by azeotropic distillation with 2-ethoxyethanol; then the alcohol is washed out with water and, after drying, the ethylcyclohexane is redistilled. The dried material has been repeatedly fractionated over Na. [Baker & Groves *J Chem Soc* 1148 1939, *Beilstein* **5** H 35, **5** III 90, **5** IV 115.]

**Ethyl cyclohexanecarboxylate** [3289-28-9] **M 156.2, b 76-77°/10mm, 92-93°/34mm, 196-196.2°/760mm,  $d_4^{20}$  0.955,  $n_D^{20}$  1.441.** Wash the ester with N sodium hydroxide solution, then water, dry with  $\text{Na}_2\text{SO}_4$  and distil it. The *amide* has **m 185-186°**. [Adkins & Cramer *J Am Chem Soc* **52** 4355 1930, Newman & Walborsky *J Am Chem Soc* **72** 4296 1950, *Beilstein* **9** III 17, **9** IV 18.]

**Ethyl 2-oxocyclohexanecarboxylate (2-ethoxycarbonylcyclohexanone, 2-carbethoxycyclohexanone, ethyl 2-ketohexahydrobenzoate)** [1655-07-8] **M 170.2, b 86-88°/3.2mm, 106°/11mm,  $d^{25}$  1.064,  $n_D^{20}$  1.47940,  $pK^{25}$  10.94 (12.87).** The ester is obtained by the decarbonylation of *ethyl 2-ketocyclohexylglyoxalate* (which in turn is prepared from cyclohexanone and diethyl oxalate in the presence of NaOEt). The *ethyl glyoxalate* (~250-265g, boiling at 105° to 165°/10-15mm) is mixed with Fe powder (1-3mg) and finely ground soft glass 0.5-1.0g in a Claisen flask (~500ml), and heated in a 40mm vacuum (bath temperature at 165-175°, not higher, to avoid unreacted ester from distilling) while CO [TOXIC] evolves and decarbonylated ester distils between 125-140°, requiring 1.5 to 2 hours for pyrolysis. The desired ester obtained (200-210g, 85.8%) has  $n_D^{25}$  1.476 to 1.479. [Snyder et al. *Org Synth Coll Vol II* 531 1943.]

The ester has FT-IR (neat) with  $\nu_{\text{max}}$  at 2939.5 (CH), 1716.2 (C=O), 1658.3 (C=O ester), 1442.5 (OH enol ?), 1365.2, 1299.3, 1219.0, 1082.9 and 832.9  $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_2$ , TMS) with  $\delta$  at 1.30 (t, 3H, ester  $\text{CH}_3$ ), 1.65 (m, ~4H, 2-methylenes), 2.25 (m, 4H 2-methylenes), 3.35 (t, 1H, partly enolised), 4.20 (q, 2H, ester  $\text{CH}_2$ ) and ~12.0 (s, enolic OH); and  $^{13}\text{C}$  NMR (15MHz,  $\text{CDCl}_3$ , TMS) with  $\delta$  at 14.17, 14.33, 21.97, 22.42, 23.31, 27.13, 29.10, 29.98, 41.55, 57.22, 60.12, 61.04, 97.72, 169.93, 171.93, 172.71, 206.19. [*Beilstein* **10** IV 2606.]

The *enol* content of *ethyl 2-oxocyclohexanecarboxylate* is higher than that of the *ethyl 2-oxocyclopentanecarboxylate* below ([611-10-9]) and varies with polarity of solvent, *viz*: 67% aqueous MeOH (22.2%), MeOH (45.1%), EtOH (61.7%),  $\text{CHCl}_3$  (54.5%),  $^*\text{C}_6\text{H}_6$  (74.8%) [Kabachnik et al. *Tetrahedron* **1** 317

1957, Gero *J Org Chem* **19** 1960 1954, see also Schreck *J Am Chem Soc* **71** 1881 1949, Buu-Hoi & Cagniant *Bull Soc Chim Fr* **10** [5] 251 1943, Lewin *Izv Akad S.S.S.R Ser fiz* **11** 413 1947, *Chem Abstr* **42** 3261 1948]. [*Beilstein* **10** H 601, **10** II 420, **10** III 2813, **10** IV 2606.]

**Ethyl 2-oxocyclopentanecarboxylate (2-ethoxycarbonylcyclopentanone, 2-carbethoxycyclopentanone, Dieckmann Ester)** [611-10-9,  $\pm$  53229-92-8] **M 156.1**, **b 79-81°/3mm**, **86-87°/6mm**, **102°/11mm**, **108-111°/15mm**. The ester is prepared in a dry N<sub>2</sub> atmosphere by adding ethyl adipate (202g, 1mole), during 2 hours, to a stirred (important to use a Hershberg stirrer, Hershberg *Ind Eng Chem, Anal Ed* **8** 313 1936) suspension of Na metal (23g, 1g.atom) in dry toluene (250ml). The reaction starts immediately and the temperature is maintained at 100-115° (oil bath) during the 2 hours and for 5 hours longer, while dry toluene is added (~750ml to 1L) so as to keep the mixture fluid and efficiently stirred, and to avoid 'caking'. The mixture is cooled to 0° and slowly added with stirring to 10% aqueous KOH (below 1°), cold H<sub>2</sub>O being added to keep the potassium salt in solution; the toluene layer is separated, washed with cold H<sub>2</sub>O (2 x 150ml), cold 10% aqueous KOH (adding cold H<sub>2</sub>O to dissolve separated potassium salt), the yellow toluene solution is finally washed with cold H<sub>2</sub>O (2 x 150ml). The combined aqueous layers, after extraction with Et<sub>2</sub>O (250ml), are run slowly, with stirring at 0°, into cold 10% AcOH. The oily ester which separates is extracted into Et<sub>2</sub>O (400ml), the aqueous layer is extracted with Et<sub>2</sub>O (4 x 250ml). The combined Et<sub>2</sub>O extracts are washed with 7% aqueous Na<sub>2</sub>CO<sub>3</sub>, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated and the residue is distilled (**b 79-81°/4mm**) to give the *keto-ester* (100-115g, 64-74%) free from ethyl adipate. An alternative synthesis in \*C<sub>6</sub>H<sub>6</sub> with slight modification provided 79-82% yield of ester. [Pinkney *Org Synth Coll Vol II* 116 1943, Cornubert & Borrel *Bull Soc Chim Fr* **47** 301 1930.]

The *enol* content of *ethyl 2-oxocyclopentanecarboxylate* is low and varies little with polarity of solvent, viz: MeOH (4.0%), EtOH (3.9%), CHCl<sub>3</sub> (3.7%), \*C<sub>6</sub>H<sub>6</sub> (8.2%) [Kabachnik et al. *Tetrahedron* **1** 317 1957, Lewin *Izv Akad S.S.S.R Ser fiz* **11** 413 1947, *Chem Abstr* **42** 3261 1948]. The ester has FT-IR (neat) with  $\nu_{\max}$  at 2979.9 (CH), 1725.6 (C=O), 1678.8 (C=O ester), 1454.5 (OH enol ?), 1370.0, 1255.0, 1190.7, 1111.9 and 1026.3 cm<sup>-1</sup>; <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 1.29 (t, *J* = ~7Hz, 3H, ester Me), 1.89 (m, 1H, C-3 Heq), 2.24 (m, 1H, C-3 Hax), 2.31 (m, 4H, C-4,5 methylenes), 3.15 (t, 1H, *J* = ~8Hz, C-1 H), 4.20 (q, *J* = ~7Hz, ester CH<sub>2</sub>); <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 14.18, 20.98, 27.43, 38.04, 54.77, 61.26, 169.37 and 212.25.

**N-Phenyl 2-oxocyclopentanecarboxamide** [4874-65-1] **M 203.2**, **m 104°** is obtained by boiling aniline (25mmol) with ethyl 2-oxocyclopentanecarboxylate (100mmol) and pyridine (0.5ml) for 2 minutes, cooling and the solid that separates is filtered off, washed with cold EtOH, then with 4% aqueous NaOH to free it from any anil formed. The solution is filtered and acidified with dilute AcOH, the precipitate is filtered off, dried and the *anilide* is recrystallised from EtOH. [Barany & Pianka *J Chem Soc* 1420 1947.] It is soluble hot EtOH and in Me<sub>2</sub>CO. A 0.3% solution of the anilide in EtOH is used for the gravimetric determination of Be, Hg V and U [Chaudhuri & Das *Anal Chim Acta* **57** 193 1971].

**1-Ethynyl-1-cyclohexanol** [78-27-3] **M 124.2**, **m 30-33°**, **32-33°**, **b 74°/12mm**, **76-78°/17mm**, **171-172°/694mm**, **180°/atm**,  $d_4^{25}$  **0.9734**,  $n_D^{25}$  **1.4801**. Dissolve it in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dilute NaHCO<sub>3</sub>, H<sub>2</sub>O again, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and distil the residue. The IR (CCl<sub>4</sub>) has  $\nu_{\max}$  at 3448 (OH), 2941 (CH), 1449-1123 and 956 cm<sup>-1</sup>, and the <sup>1</sup>H NMR (CCl<sub>4</sub>) has  $\delta$  at 3.2 (OH), 2.5 ( $\equiv$ CH), 1.70 (m 10H, CH<sub>2</sub>) [Hasbrouck & Kiessling *J Org Chem* **38** 2103 1972]. [*Beilstein* **6** II 100.] **TOXIC**.

**Eucaliptol (1,8-cineol, 1,8-epoxy-*p*-menthane, 1,3,3-trimethyl-2-oxabicyclo[2.2.2]-octane)** [470-82-6] **M 154.2**, **m 1.3°**, **1.5°**, **b 39-39.3°/4mm**, **176-176.4°/760mm**,  $d_4^{20}$  **0.9232**,  $n_D^{20}$  **1.4575**. Purify 1,8-cineol by dilution with an equal volume of petroleum ether, then saturate it with dry HBr. The precipitate is filtered off, washed with small volumes of petroleum ether, then cineole is regenerated by stirring the crystals with H<sub>2</sub>O. It can also be purified *via* its *o*-cresol or resorcinol addition compounds. Store it over Na until required. Purify it also by fractional distillation. It is insoluble in H<sub>2</sub>O but soluble in organic solvents. [IR: Kome et al. *Nippon Kagaku Zasshi [J Chem Soc Japan (Pure Chem Sect)]* **80** 66 1959, *Chem Abstr* 603 1961, *Beilstein* **17** II 32, **17/1 V 273**.]

**(+)- $\alpha$ -Fenchol (1R-1,3,3-trimethylnorbornan-2-ol)** [1632-73-1] **M 154.3**, **m 40-43°**, **47-47.5°**, **b 201-**

**202°**,  $[\alpha]_{\text{D}}^{20} +12.5^{\circ}$  (c 10, EtOH). It is prepared by reduction of (–)-fenchone and is purified by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, or distillation, or both. The 2-carboxybenzoyl (monophthalate) derivative has m 146.5-147.5°  $[\alpha]_{\text{D}}^{20} -20.4^{\circ}$  (EtOH), and the 2-phenylurethane has m 81°. [Beckmann & Metzger *Chem Ber* 89 2738 1956]. [Beilstein 6 III 288, 6 IV 278.]

(+)- Fenchone (1*S*-1,3,3-trimethylnorbornan-2-one) [4695-62-9] M 152.2, m 5-7°, 6.1°, b 63-65°/13mm, 66°/15mm, 122°/10mm,  $d_4^{20}$  0.9434,  $n_{\text{D}}^{20}$  1.4636,  $[\alpha]_{\text{D}}^{20} +66.9^{\circ}$  (neat, or in c 1.5, EtOH),  $[\alpha]_{546}^{20} +60.4^{\circ}$  (neat). The oily liquid is purified by distillation in a vacuum and is very soluble in EtOH and Et<sub>2</sub>O. [Boyle et al. *J Chem Soc, Chem Commun* 395 1971, Hüchel *Justus Liebigs Ann Chem* 549 186 1941, (±)-isomer: Braun & Jacob *Chem Ber* 66 1461 1933.] It forms two oximes, the *cis*-oxime has m 167° (crystallises from petroleum ether)  $[\alpha]_{\text{D}}^{20} +46.5^{\circ}$  (c 2, EtOH), the *O*-benzoyloxime has m 81°,  $[\alpha]_{\text{D}}^{18} +49^{\circ}$  (EtOH), and the oxime-HCl has m 136°(dec). The *trans*-oxime has m 123° (from petroleum ether)  $[\alpha]_{\text{D}}^{18} +148^{\circ}$  (c 2, EtOH) and the *O*-benzoyloxime has m 125°  $[\alpha]_{\text{D}}^{20} +128.5^{\circ}$  (c 2, EtOH) [Hüchel *Justus Liebigs Ann Chem* 549 186 1941, Hüchel & Sachs *Justus Liebigs Ann Chem* 498 166 1932]. [Beilstein 7 III 212, 7 IV 212.]

(–)-Fenchone (1*R*-1,3,3-trimethylnorbornan-2-one) [7787-20-4] M 152.2, m 5.2°, b 67.2°/10mm, 191-195°/atm,  $d_4^{20}$  0.9484,  $n_{\text{D}}^{20}$  1.4630,  $[\alpha]_{\text{D}}^{20} -66.8^{\circ}$  (neat). Purification is as for the (+)-enantiomer above and should have the same physical properties except for opposite optical rotations. UV has  $\lambda_{\text{max}}$  285nm ( $\epsilon$  12.29). [Braun & Jacob *Chem Ber* 66 1461 1933, UV: Ohloff et al. *Chem Ber* 90 106 1957.] [Beilstein 7 III 392, 7 IV 212.]

**Gibberillic acid A<sub>3</sub>** (gibberillin A<sub>3</sub>) [77-06-5] M 346.4, m 233-235°(dec),  $[\alpha]_{546}^{20} +92^{\circ}$  (c 1, MeOH),  $[\alpha]_{\text{D}}^{20} +93^{\circ}$  (c 0.5, MeOH), pK 4.0. It crystallises from EtOAc, EtOAc/petroleum ether, MeOH/petroleum ether or Me<sub>2</sub>CO/petroleum ether. [Cross *J Chem Soc* 3022 1960, Beilstein 18 III/IV 6533.]

**1,2,3,4,5,6-Hexachlorocyclohexane** [ $\alpha$ -319-84-6,  $\gamma$ -58-89-9] M 290.8, m 158° ( $\alpha$ -), 312° ( $\beta$ -), 112.5° ( $\gamma$ -isomer). Crystallise it from EtOH. Purify it also by zone melting. **Possible CANCER AGENT, TOXIC.** [ $\alpha$ : Beilstein 1 H 23,  $\gamma$ : Beilstein 5 I 8, many isomers : Beilstein 5 III 41, 5 IV 55.]

**1,2,3,4,5,5-Hexachlorocyclopenta-1,3-diene** [77-47-4] M 272.8, b 80°/1mm, 83-84°/3mm, 234°/atm,  $d_4^{25}$  1.702,  $n_{\text{D}}^{25}$  1.5628. Dry the diene with MgSO<sub>4</sub>, filter, and distil it under vacuum in a nitrogen atmosphere. Irritates skin and eyes, **HIGHLY TOXIC.** [McBee et al. *J Am Chem Soc* 77 4378 1955, UV spectra: Idol et al. *J Org Chem* 20 1746 1955, Beilstein 5 III 308, 5 IV 381.]

**Hexahydromandelic acid** [*R*(–)- 53585-93-6, *S*(+) 61475-31-8] M 158.2, m 127-129°, 128-129°, 129.7°,  $[\alpha]_{\text{D}}^{20}$  (–) and (+) 25.5° (c 1, AcOH) and  $[\alpha]_{\text{D}}^{20}$  (–) and (+) 13.6° (c 7.6, EtOH). It forms hexagonal clusters on recrystallisation from CCl<sub>4</sub> or Et<sub>2</sub>O. [Wood & Comley *J Chem Soc* 2638 1924, Lettré et al. *Chem Ber* 69 1594 1936]. The *racemate* has m 137.2-137.6° (134-135°) [Smith et al. *J Am Chem Soc* 71 3772 1949]. [Beilstein *R*- 10 II 5; *S*- 10 II 6.]

**Hexamethyl(Dewar)benzene (HMDB, 1,2,3,4,5,6-hexamethyl-bicyclo[2.2.0]hexa-2,5-diene)** [7641-77-2] M 162.3, m 7.5°, b 60°/20mm, ~152°/760mm,  $d_4^{20}$  0.8125,  $n_{\text{D}}^{20}$  1.4480. HMDB is obtained in ~80% yield when 2-butyne (dimethylacetylene, flammable gas b 27°/atm, [503-17-3]) in \*C<sub>6</sub>H<sub>6</sub> or CH<sub>2</sub>Cl<sub>2</sub> containing anhydrous AlCl<sub>3</sub> is stirred at 35° for 5-7 hours. The brown reaction mixture is poured onto crushed ice, washed with dilute aqueous NaOH, and HMDB is isolated from the organic phase by fractional vacuum distillation. Hexamethylbenzene (m 165-166°, [87-85-4]) is a by-product (~12-18%) together with a mixture (~2%) of *syn*- (m 127°) and *anti*- (m 196°) octamethyltricyclo[4.2.0.0<sup>2,5</sup>]octa-3,7-diene, and octamethylcyclooctatetraene (m 113°). Although longer reaction times increase the conversion of 2-butyne, they lower the yields of HMDB in favour of the other by-products. It is fairly thermally stable with half-life time conversions to hexamethylbenzene with 105 hours/120°, 5.5 hours/140°, and 2.1 hours/150°. It is best to store it at low temperature and away from light as radiation from a low pressure UV lamp converts it mainly to hexamethylbenzene with ~20-25% of the valence isomer hexamethylprismane (m 91°). The UV spectrum

exhibits tail end absorption from 220 to 250nm; the FT-IR (neat) has  $\nu_{\max}$  at 2950.0, 1683.5, 1439.5, 1368.8, 1275.4, 1222.2, 1064.8, 736.2 and 490.7  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.08 (s, 6H, 1,4-( $\text{CH}_3$ )<sub>2</sub>) and 1.69 (s, 12H, 2,3,5,6-( $\text{CH}_3$ )<sub>4</sub>), and the  $^{13}\text{C}$  NMR (15MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 10.06 (1,4-Me carbons) and 11.22 (2,3,5,6-Me carbons), 55.68 (2 saturated carbons) and 144.09 (4 olefinic carbons). [Schäfer & Hellmann *Angew Chem Internat Edn* **6** 518 1967.] Purify it also by passing it neat through alumina or in  $^*\text{C}_6\text{H}_6$  or  $\text{CH}_2\text{Cl}_2$  solution. [Traylor & Miksztal *J Am Chem Soc* **109** 2770 1987].

**Humulon** [26472-41-3] **M 362.5, m 65-66.5°,  $[\alpha]_{\text{D}}^{26} -212^\circ$  (95% EtOH)**. Crystallise humulon from  $\text{Et}_2\text{O}$ . It dissolves slightly in hot  $\text{H}_2\text{O}$  but precipitates on cooling. It has  $\lambda_{\max}$  nm ( $\epsilon$ ) at 237 (13,760) and 282 (8,330) in EtOH. [Wollmer *Chem Ber* **49** 780 1916, Carson *J Am Chem Soc* **73** 4652 1951, *Beilstein* **8** II 537, **8** III 4034, **8** IV 3410.]

**1-Hydroxymethyladamantane** [770-71-8] **M 166.3, m 115°**. Dissolve the adamantane in  $\text{Et}_2\text{O}$ , wash it with aqueous 0.1N NaOH and  $\text{H}_2\text{O}$ , dry over  $\text{CaCl}_2$ , evaporate and recrystallise the residue from aqueous MeOH. [Stetter et al. *Chem Ber* **92** 1629 1959, *Beilstein* **6** IV 400.]

**N-Hydroxy-5-norbornene-2,3-dicarboxylic acid imide** [21715-90-2] **M 179.2, m 165-166°, 166-169°,  $\text{pK}_{\text{Est}} \sim 6$** . Dissolve the imide in  $\text{CHCl}_3$ , filter, evaporate and recrystallise from EtOAc. The IR (nujol) has  $\nu_{\max}$  1695, 1710 and 1770 (C=O), and 3100 (OH)  $\text{cm}^{-1}$ . The *O*-acetyl derivative has **m** 113-114° (from EtOH) with IR bands at  $\nu_{\max}$  1730, 1770 and 1815  $\text{cm}^{-1}$  only, and the *O*-benzoyl derivative has **m** 143-144° (from propan-2-ol or  $^*\text{C}_6\text{H}_6$ ). [Bauer & Miarka *J Org Chem* **24** 1293 1959, Fujino et al. *Chem Pharm Bull Jpn* **22** 1857 1974]. [*Beilstein* **21/10** V 188.]

***i*-Inositol (*myo*)** See in "Miscellaneous", Chapter 7.

**$\alpha$ -Ionone (*trans*-+)** [127-41-3] **M 192.3, b 86-87°/1.9mm, 131°/13mm,  $d_4^{20}$  0.929,  $n_{\text{D}}^{20}$  1.5497,  $[\alpha]_{\text{D}}^{20} +401^\circ$  (neat)  $+415^\circ$  (EtOH)**. Purify  $\alpha$ -ionone through a spinning band fractionating column. The semicarbazone has **m** 157-157.5° (from EtOH) and  $[\alpha]_{\text{D}}^{20} +433^\circ$  (c 4,  $^*\text{C}_6\text{H}_6$ ). [Naves *Helv Chim Acta* **30** 769 1947, CD: Ohloff et al. *Helv Chim Acta* **56** 1874 1973, Buchacker et al. *Helv Chim Acta* **56** 2548 1973, *Beilstein* **7** H 168, **7** III 640, **7** IV 363.]

**$\beta$ -Ionone** [79-77-6] **M 192.3, b 150-151°/24mm,  $d_4^{20}$  0.945,  $n_{\text{D}}^{20}$  1.5211,  $\epsilon_{296\text{nm}}$  10,700**. Convert  $\beta$ -ionone to the semicarbazone (**m** 149°) by adding semicarbazide hydrochloride (50g) and potassium acetate (44g) in water (150ml) to a solution of  $\beta$ -ionone (85g) in EtOH. (More EtOH is added to redissolve any  $\beta$ -ionone that precipitates.) The semicarbazone crystallises on cooling in an ice-bath and is recrystallised from EtOH or 75% MeOH to constant **m** (148-149°). The semicarbazone (5g) is shaken at room temperature for several days with petroleum ether (20ml) and M  $\text{H}_2\text{SO}_4$  (48ml); then the petroleum ether layer is washed with water and dilute aqueous  $\text{NaHCO}_3$ , dried and the solvent is evaporated. The  $\beta$ -ionone is distilled under vacuum. (The customary steam distillation of  $\beta$ -ionone semicarbazone did not increase the purity.) [Young et al. *J Am Chem Soc* **66** 855 1944]. [*Beilstein* **7** H 167, **7** I 109, **7** II 140, **7** III 634, **7** IV 361.]

**( $\pm$ )-Irone (6-methylionone,  $\pm$ -*trans*-( $\alpha$ )-4*t*-[2,5,6,6-tetramethylcyclohex-2-yl]but-3*t*-en-2-one)** [79-69-6] **M 206.3, b 85-86°/0.05mm, 109°/0.7mm,  $d_4^{20}$  0.9340,  $n_{\text{D}}^{20}$  1.4998**. If large amounts are available, then fractionate through a Podbielniak column or an efficient spinning band column, but small amounts are distilled using a Kügelrohr apparatus. The 4-phenylsemicarbazone has **m** 174-175° (165-165.5°). [IR: Seidel & Ruzicka *Helv Chim Acta* **35** 1826 1952, Naves *Helv Chim Acta* **31** 1280 1948, Lecomte & Naves *J Chim Phys* **53** 462 1956, *Beilstein* **7** IV 378.]

***dl*-Isoborneol** [124-76-5] **M 154.3, m 212° (sealed tube)**. Crystallise isoborneol from EtOH or petroleum ether (b 60-80°). It sublimes in a vacuum. The 4-nitrobenzoyl derivative has **m** 153°. [Yager & Morgan *J Am Chem Soc* **57** 2081 1935, *Beilstein* **6** II 80, **6** III 299, **6** IV 281.]

**(-)- $\gamma$ -Isocaryophyllene (1*R*,9*S*-8-methylene-4,11,11-trimethylbicyclo[7.2.0]undec-4-ene)** [118-65-0] **M 204.4, b 122-124°/12mm, 131-133°/16mm, 130-131°/24mm, 271-273°/atm,  $d_4^{20}$  0.8959,  $n_{\text{D}}^{20}$  1.496,  $[\alpha]_{546}^{20}$**

$-31^\circ$ ,  $[\alpha]_D^{20} -27^\circ$  (neat). Purify it by vacuum distillation or GLC using a nitrile-silicone column [Corey et al. *J Am Chem Soc* **86** 485 1964, Ramage & Simonsen *J Chem Soc* 741 1936, Kumar et al. *Synthesis* 461 1976]. [Beilstein **5** III 1085.]

(-)- $\beta$ -Isolongifolene (1-*R*-(-)-2,2,7,7-tetramethyltricyclo[6.2.1.0<sup>1,6</sup>]undec-5-ene) [1135-66-6] **M 204.4, b 82-83<sup>o</sup>/0.4mm, 144-146<sup>o</sup>/30mm, 255-256<sup>o</sup>/atm,  $d_4^{20}$  0.930,  $n_D^{20}$  1.4992,  $[\alpha]_{546}^{20} -166^\circ$ ,  $[\alpha]_D^{20} -38^\circ$  (c 1, EtOH). Reflux it over, and distil it from Na. [Zeiss & Arakawa *J Am Chem Soc* **76** 1653 1954, IR: Reinaecker & Graafe *Angew Chem, Int Ed Engl* **97** 348 1985, UV and NMR: Ranganathan et al. *Tetrahedron* **26** 621 1970, Beilstein **5** IV 1191.]**

**Isophorone** [78-59-1] **M 138.2, b 94<sup>o</sup>/16mm,  $d_4^{20}$  0.921,  $n_D^{20}$  1.4778**. Wash isophorone with aqueous 5% Na<sub>2</sub>CO<sub>3</sub> and then distil it under reduced pressure immediately before use. Alternatively, it can be purified via the semicarbazone. [Erskine & Waight *J Chem Soc* 3425 1960, Beilstein **7** IV 165.]

**Isopinocampheol (pinan-3-ol, 2,6,6-trimethylbicyclo[3.1.1]heptan-3-ol)** [1*S*,2*S*,3*S*,5*R*-(+)- 27779-29-9, 1*R*,2*R*,3*R*,5*S*-(-)- 25465-65-0] **M 154.25, m 52-55<sup>o</sup>, 55-56<sup>o</sup>, 55-57<sup>o</sup>, b 103<sup>o</sup>/11mm,  $n_D^{20}$  1.4832,  $[\alpha]_{546}^{20}$  (+) and (-) 43<sup>o</sup>.  $[\alpha]_D^{20}$  (+) and (-) 36<sup>o</sup> (c 20, EtOH)**. Dissolve it in Et<sub>2</sub>O, dry it over MgSO<sub>4</sub>, filter, evaporate, then recrystallise it from petroleum ether. Also recrystallise it from aqueous EtOH and distil it in a vacuum. [Kergomard & Geneix *Bull Soc Chim Fr* 394 1958, Zweifel & Brown *J Am Chem Soc* **86** 393 1964.] The 3,4-dinitrobenzoyl derivative has **m 100-101<sup>o</sup>**, the phenylcarbamoyl derivative has **m 137-138<sup>o</sup>** and the acid-phthalate has **m 125-126<sup>o</sup>**. [Beilstein **6** III 282, 283.]

**Isopropenylcyclobutane** [3019-22-5] **M 98.1, b 98.7<sup>o</sup>/760mm,  $d_4^{20}$  0.7743,  $n_D^{20}$  1.438**. Purify the cyclobutane by preparative chromatography (silicon oil column), or fractional distillation. Dry it over molecular sieves. Its IR (film) has  $\nu_{\max}$  at 1640 (C=C), 887 and 1773 (C-H) cm<sup>-1</sup>. [Chiurdohlu & Van Walle *Bull Soc Chim Belg* **66** 612 1957, Beilstein **5** IV 255.]

**(1*S*,2*S*,4*R*)-(+)-Limonene-1,2-diol** [1*S*,2*S*,4*R*)-(+)-4-isopropenyl-1-methylcyclohexan-1,2-diol, (1*S*,2*S*,4*R*)-(+)-*p*-menth-8-en-1,2-diol, (+)-1-hydroxyneodihydrocarveol, neolimonene glycol] [38630-75-0] **M 170.3, m 68-72<sup>o</sup>, 70-70.5<sup>o</sup>, 72-73<sup>o</sup>,  $[\alpha]_D^{25} +55^\circ$  (c 1, Me<sub>2</sub>CO, also +45<sup>o</sup> and +53.7<sup>o</sup> were reported)**. It has been prepared from commercial grade (+)-limonene 1,2-oxide (500g,  $[\alpha]_D^{27} +63.3^\circ$  (neat) by stirring at 0<sup>o</sup> with 6% H<sub>2</sub>SO<sub>4</sub> (2.5L) for 5 hours, the solution is filtered, and the crude diol hydrate is dissolved in hot CHCl<sub>3</sub> separated from H<sub>2</sub>O, concentrated, and crystallised from CHCl<sub>3</sub> or petroleum ether to give the anhydrous diol (375g, 67%) **m 70-70.5<sup>o</sup>**. The tri-hydrate crystallises from H<sub>2</sub>O in plates with **m 60<sup>o</sup>**. On hydrogenation in EtOAc over Pt<sub>2</sub>O, it provides (+)-1-hydroxyneocarvomenthol, **m 88<sup>o</sup>,  $[\alpha]_D^{25} +48^\circ$  (Me<sub>2</sub>CO)** after recrystallisation from \*C<sub>6</sub>H<sub>6</sub>-petroleum ether (b 70-110<sup>o</sup>). [Royals & Leffingwell *J Am Chem Soc* **31** 1937 1966, Newhall *J Org Chem* **29** 185 1964, Beilstein **6** H 753, **6** II 758, **6** III 4137, **6** IV 5294.]

**Lupulon ( $\beta$ -lupulic acid, bitter acid)** [468-28-0] **M 414.6, m 92-94<sup>o</sup>,  $pK_{\text{Est}(1)} \sim 4.2$ ,  $pK_{\text{Est}(2)} \sim 9.7$** . Recrystallise Lupulon from 90% MeOH, hexane or petroleum ether at low temperature. It has been purified by chromatography through Kieselgel. [Wieland et al. *Chem Ber* **102** 2012 1925, Riedl *Chem Ber* **85** 692 1952, Beilstein **7** II 856, **7** III 4752, **7** IV 2866.]

***l*-(-)-Menthone (2*S*,5*R*-2-isopropyl-5-methylcyclohexan-1-one)** [14073-97-3] **M 154.3, b 98-100<sup>o</sup>/14mm, 205-208<sup>o</sup>/atm, 207-210<sup>o</sup>/760mm,  $d_4^{25}$  0.8930,  $n_D^{20}$  1.4505,  $[\alpha]_D^{20} -28^\circ$  (neat)**. It is obtained by adding pure *l*-(-)-menthol (4.5g, see [2216-51-5]) in four portions to chromic acid [prepared from Na<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub>·2H<sub>2</sub>O (60g, CARE CARCINOGENIC) in H<sub>2</sub>O (300ml) and concentrated H<sub>2</sub>SO<sub>4</sub> (27ml) with cooling) when the temperature rises, but should be kept at 55<sup>o</sup> (warm if necessary). A black spongy mass is first formed which becomes soft and melts to a dark brown oil on the surface as the temperature rises. The temperature drops when the oxidation is complete, the mixture is cooled and extracted with Et<sub>2</sub>O (100ml), the extract is separated, washed with 5% aqueous NaOH (100ml, CARE) several times (~ x 3) until the colour is light yellow, then H<sub>2</sub>O and dried (Na<sub>2</sub>SO<sub>4</sub> or Mg<sub>2</sub>SO<sub>4</sub>). The ethereal solution is evaporated and the residual oil is redistilled at

atmospheric pressure or in a vacuum through a short column to give pure *l*-menthone (38g, 91%). [Brown & Garg *J Am Chem Soc* **83** 2952 1961, Hussey & Baker *J Org Chem* **25** 1434 1960.] This menthone is the most frequently naturally occurring of the four optically active isomers. It has been used for the optical resolution of diols *via* acetalisation [Harada & Oku *Synlett* 95 1994], and the menthone spirocyclic 1,3-dioxane-4,6-diones have been used for asymmetric [2 + 2]cycloaddition and Diels-Alder reactions [Sato et al. *Tetrahedron* **47** 7271 1991]. The *oxime* has **m 59°** (from petroleum ether), the *semicarbazone* has **m 189-191°** (from MeOH, EtOH or aqueous EtOH), the *phenylhydrazone* has **m 53°** (from aqueous EtOH), and the *2,4-dinitrophenylhydrazone* has **m 146°** (from EtOH, aqueous EtOH or EtOH/EtOAc). [*Beilstein* 7 H 38, 7 I 34, 7 II 39, 7 III 152, 7 IV 87.] It is used in flavourings and perfume.

**1*R*-(–)-Menthyl chloride (1*S*,2*R*,4*R*-2-chloro-1-isopropyl-4-methylcyclohexane)** [16052-42-9] **M 174.7, m –20.1° to –16.5°, b 88.5°/12.5mm, 101-105°/21mm,  $d_4^{20}$  0.936,  $n_D^{20}$  1.463,  $[\alpha]_D^{20}$  –52.4° (neat)**. Dissolve menthyl chloride in petroleum ether (b 40-60°), wash it with H<sub>2</sub>O, conc H<sub>2</sub>SO<sub>4</sub> until no discoloration of the organic layer occurs (care with conc H<sub>2</sub>SO<sub>4</sub> during shaking in a separating funnel), again with H<sub>2</sub>O and dry it (MgSO<sub>4</sub>). Evaporate, and distil the residual oil through a Claisen head with a Vigreux neck of *ca* 40 cm length. [Smith & Wright *J Org Chem* **17** 1116 1952, Barton et al. *J Chem Soc* 453 1952, *Beilstein* 5 III 134, 5 IV 152.]

**1-Methyladamantane** [768-91-2] **M 150.2, m 103°, 104°**. Purify it by zone melting, chromatography through an Al<sub>2</sub>O<sub>3</sub> column and eluting with pentane, and sublime it repeatedly at 90-95°/12mm. [Stetter et al. *Chem Ber* **92** 1629 1959, Schleyer & Nicholas *Tetrahedron Lett* **9** 305 1961, *Beilstein* 5 IV 479.]

**2-Methyladamantane** [700-56-1] **M 150.2, 144-146°**. Purify it by zone melting, chromatography through an Al<sub>2</sub>O<sub>3</sub> column and eluting with pentane. Recrystallise it from EtOH and sublime it repeatedly at 90-95°/12mm. [Schleyer & Nicholas *J Am Chem Soc* **83** 182 1961, Molle et al. *Can J Chem* **65** 2428 1987.]

**Methylcyclohexane** [108-87-2] **M 98.2, b 100.9°,  $d_4^{25}$  0.7650,  $n_D^{20}$  1.4231,  $n_D^{25}$  1.42058**. Passage through a column of activated silica gel gives material transparent down to 220nm. It can also be purified by passage through a column of activated basic alumina, or by azeotropic distillation with MeOH, followed by washing out the MeOH with H<sub>2</sub>O, drying and distilling. Methylcyclohexane can be dried with CaSO<sub>4</sub>, CaH<sub>2</sub> or sodium. It has also been purified by shaking with a mixture of conc H<sub>2</sub>SO<sub>4</sub> and HNO<sub>3</sub> in the cold, washing with H<sub>2</sub>O, drying with CaSO<sub>4</sub> and fractionally distilling it from potassium. Percolation through a column of Celite impregnated with 2,4-dinitrophenylhydrazine (DNPH), phosphoric acid and H<sub>2</sub>O (prepared by grinding 0.5g DNPH with 6ml 85% H<sub>3</sub>PO<sub>4</sub>, then mixing with 4ml of distilled H<sub>2</sub>O and 10g of Celite) removes carbonyl-containing impurities. [Cowan et al. *J Chem Soc* 1865 1939, *Beilstein* 5 III 65, 5 IV 94.]

***cis*- and *trans*-2-Methylcyclohexanol** [583-59-5] **M 114.2, b 65°/20mm, 167.6°/760mm,  $d_4^{20}$  0.922,  $n_D^{20}$  1.46085**. Dry 2-methylcyclohexanol with Na<sub>2</sub>SO<sub>4</sub> and fractionate it under vacuum. Note: The *cis*-isomer has **b 165°/760mm**, and the *trans*-isomer has **b 166.5°/760mm**. [Eliel & Haber *J Org Chem* **23** 2041 1958, *Beilstein* 6 III 61, 6 IV 100.]

***cis*- and *trans*-3-Methylcyclohexanol** [591-23-1] **M 114.2, b 69°/16mm, 172°/760mm,  $d_4^{20}$  0.930,  $n_D^{20}$  1.45757,  $n_D^{25.5}$  1.45444**. Dry 3-methylcyclohexanol with Na<sub>2</sub>SO<sub>4</sub> and fractionate it under vacuum. Note: The *cis*-isomer has **b 173°/760mm**, and the *trans*-isomer has **b 168-169°/760mm**. [Eliel & Haber *J Org Chem* **23** 2041 1958, *Beilstein* 6 IV 102.]

**4-Methylcyclohexanone** [589-92-4] **M 112.2, m –40.6°, b 68°/23mm, 165.5°/743mm,  $d_4^{20}$  0.914,  $n_D^{20}$  1.44506**. Dry the ketone with CaSO<sub>4</sub>, then fractionally distil it. The *semicarbazone* has **m 197°, 203.5°(dec)** (from MeOH or EtOH). [White & Bishop *J Am Chem Soc* **62** 8 1945, Vogel & Oommen *J Chem Soc* 774 1930, *Beilstein* 7 III 63, 7 IV 44.]

**1-Methylcyclohexene** [591-49-1] **M 96.2, m –120.4°, b 107.4-108°/atm, 110-111°/760mm,  $d_4^{20}$  0.813,  $n_D^{20}$  1.451**. Free it from hydroperoxides by passing through a column containing basic alumina or refluxing with cupric stearate, filter and fractionally distil it from sodium. [Vogel *J Chem Soc* 1332 1938, Cope et al. *J Am*

*Chem Soc* 79 4729 1957, *Beilstein* 5 III 197, 5 VI 245.]

**Methylcyclopentane** [96-37-7] **M 84.2, b 64.32°/400mm, 71.8°/atm,  $d_4^{20}$  0.749,  $n_D^{20}$  1.40970,  $n_D^{25}$  1.40700.** Purification procedures include passage through columns of silica gel (prepared by heating in nitrogen to 350° prior to use) and activated basic alumina, distillation from sodium-potassium alloy, and azeotropic distillation with MeOH, followed by washing out the methanol with water, drying and distilling. It can be stored with CaH<sub>2</sub> or sodium. [*Vogel, J Chem Soc* 1331 1938, *Beilstein* 5 III 55, 5 IV 84.]

**Methylnorbornene-2,3-dicarboxylic anhydride (5-methylnorborn-5-ene-2-endo-3-endo-dicarboxylic anhydride)** [25134-21-8] **M 178.2, m 88.5-89°.** Purify the anhydride by thin layer chromatography on Al<sub>2</sub>O<sub>3</sub> (previously boiled in EtOAc) and eluted with hexane/\*C<sub>6</sub>H<sub>6</sub> (1:2), then recrystallise it from \*C<sub>6</sub>H<sub>6</sub>/hexane. The *free acid* has **m 118.5-119.5°.** [Miranov et al. *Tetrahedron* 19 1939 1963, *Beilstein* 17/11 V 199.]

**2,5-Norbornadiene (bicyclo[2.2.1]hepta-2,5-diene, NBD)** [121-46-0] **M 92.1, b 89°,  $d_4^{20}$  0.854,  $n_D^{20}$  1.4707.** Purify the diene by distillation from activated alumina [Landis & Halpern *J Am Chem Soc* 109 1746 1987]. [*Beilstein* 5 IV 879.]

**cis-endo-5-Norbornene-2,3-dicarboxylic anhydride (carbic anhydride, 3 $\alpha$ ,4,7,7, $\alpha\alpha$ -tetrahydro-4 $\alpha$ ,7 $\alpha$ -methanoisobenzofuran-1,3-dione)** [129-64-6] **M 164.2, m 164.1°, 164-165°, 165-167°,  $d_4^{20}$  1.417.** It forms crystals from petroleum ether, hexane or cyclohexane. It is hydrolysed by H<sub>2</sub>O to form the acid [Diels & Alder *Justus Liebigs Ann Chem* 460 98 1928, Maitte *Bull Soc Chim Fr* 499 1959]. The *exo-exo-isomer* has **m 142-143° (from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether)** [Alder & Stein *Justus Liebigs Ann Chem* 504 216 1933]. [*Beilstein* 17 II 461.]

**( $\pm$ )-endo-2-Norbornylamine hydrochloride ( $\pm$  *endo*[2.2.1]hept-2-ylamine HCl)** [14370-45-7] **M 147.7, m ~295°(dec),  $pK_{Est}$  ~ 9.0(free base).** Recrystallise the salt from MeOH/EtOAc or EtOH/ Et<sub>2</sub>O. The *free base* has **m 75-80°, b 156-157°/atm** and the *picrate* has **m 179-180° (from aqueous MeOH).** [*Beilstein* 12 III 160.]

**Norbornylene (bicyclo[2.2.1]hept-2-ene)** [498-66-8] **M 94.2, m 44-46°, b 96°.** Reflux it over Na, and distil it [Gilliom & Grubbs *J Am Chem Soc* 108 733 1986]. It has also been purified by sublimation *in vacuo* onto an ice-cold finger. [Woon et al. *J Am Chem Soc* 108 7990 1986, *Beilstein* 5 IV 394.]

**( $\pm$ )-exo-2-Norbornylformate** [41498-71-9] **M 140.2, b 65-67°/16mm, 80-81°/25mm,  $d_4^{20}$  1.048,  $n_D^{20}$  1.4620.** Shake with NaHCO<sub>3</sub> and distil it *in vacuo* (*exo-borneol* has **m 124-126°**). *Alternatively*, mix the ester with formic acetic anhydride overnight and fractionate. [*Beilstein* 6 III 219.]

**Norcamphor (bicyclo[2.2.1]heptan-2-one,  $\pm$  norbornan-2-one)** [497-38-1] **M 110.2, m 94-95°, 95.5-96.5°, b 89-94°/60mm.** Crystallise it from water and sublime it *in vacuo*. It has  $\lambda_{max}$  at 287nm (EtOH). The *semicarbazone* has **m 196-196.5° (from EtOH/H<sub>2</sub>O)**. The *2,4-dinitrophenylhydrazone* has **m 137-138° (from EtOH)**. [Wildman & Hemminger *J Org Chem* 17 1641 1952, Wood & Roberts *J Org Chem* 23 1124 1957, Bixter & Niemann *J Org Chem* 23 742 1958, *Beilstein* 7 III 243, 7 IV 139.]

**1,2,3,4,5-Pentamethylcyclopentadiene (Cp')** [4045-44-7] **M 136.2, b 55-60°/13mm, 58°/13mm, 58.3°/13.5mm,  $d_4^{25}$  0.870,  $n_D^{20}$  1.4740.** Of the many syntheses of this useful ligand, the following is the most economical on materials and can be scaled up. Strictly anhydrous and anaerobic conditions should be used and reagents should be dried appropriately, degassed, and precaution against fire should be exercised. Three steps are involved:

*Step 1:* Under argon, lithium (58g, 8.36moles, with 0.02%Na, of ~3.2mm diameter cut into the flask in ~5mm lengths) is covered with Et<sub>2</sub>O (1600ml, freshly distilled from ~3:1w/w K/Na benzophenone), and 2-bromo-2-butene (120g, 0.88mole, 90.4ml of a molecular sieves 4Å dried commercial *cis-* and *trans-* mixture or prepared according to Bordwell & Landis *J Am Chem Soc* 79 1593 1957) is added slowly with stirring (10ml aliquots slowly at first until the reaction begins; with evolution of bubbles and cloudiness due to separation of LiBr), and

at such a rate as to maintain gentle reflux of the Et<sub>2</sub>O.

*Step 2:* After addition is complete (~1.5 hours), still under argon, a mixture of ethyl acetate (166g, 1.88mole, 184ml, dried over molecular sieves 4Å) and 2-bromo-2-butene (430g, 3.18mole, 324ml as above) are then added dropwise with stirring, while carefully maintaining gentle reflux over a period of 4-5 hours. When this addition is completed, a further volume of dry EtOAc (50g, 55.4ml) and a further portion of 2-bromo-2-butene (~10-20g, 7.5-15ml) are added, while stirring is continued, until refluxing of Et<sub>2</sub>O ceases. The mixture is allowed to cool over 4 hours, saturated aqueous NH<sub>4</sub>Cl solution is added dropwise to hydrolyse unreacted Li, the Et<sub>2</sub>O layer is collected, the aqueous layer is extracted with Et<sub>2</sub>O (3 x 200ml), and the combined Et<sub>2</sub>O solutions are evaporated to ~350ml.

*Step 3:* The Et<sub>2</sub>O concentrate is added, with stirring, to a slurry of *p*-toluenesulfonic acid monohydrate (26g) in Et<sub>2</sub>O (500ml) under a reflux condenser, at such a rate the the solvent refluxes gently. The mixture is then stirred for 5 minutes after refluxing ceases, and poured into saturated aqueous NaHCO<sub>3</sub> (1200ml) containing Na<sub>2</sub>CO<sub>3</sub> (19g). The yellow aqueous phase is removed, extracted with Et<sub>2</sub>O (3 x 200ml), and the combined Et<sub>2</sub>O solutions are dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated to 250-300ml (rotavap), then trap-to-trap distilled *in vacuo* (bath temperature at 35-40°) to give a yellow liquid (85% pure by GC) which is fractionally distilled in a vacuum under N<sub>2</sub> using a 50cm vigreux column. The fraction with b 65-70°/20mm collected (142g, 53% yield based on EtOAc used) as pale-yellow liquid is *1,2,3,4,5-pentamethylcyclopentadiene* (92% pure by GC). An additional fraction (15g, 5%) with b 70-75°/20mm is 85% pure by GC. The liquids are colourless to pale-yellow in colour with a sweet olefinic odour, and are pure enough for use as ligands. They should be stored in a freezer under N<sub>2</sub> or argon. In comparison with the unsubstituted *cyclopentadiene* it is a stronger donor of electron density, exerting considerably enhanced thermal stability, and the metal complexes that it forms are generally more soluble and easier to crystallise.

The *pentamethylpentadiene* has FT-IR (neat) with  $\nu_{\max}$  at 2960 (vs), 2915 (vs), 2855 (vs), 2735 (w), 1660 (m), 1640 (w), 1390 (s), 1355 (m), 1150 (w), 1105 (mw), 1048 (w), 840 (mw) and 668 (w) cm<sup>-1</sup>; <sup>1</sup>H NMR (60MHz, CCl<sub>4</sub>, TMS) with  $\delta$  at 2.4 (m, 1H), 1.75 (br s, 13H, 2,3,4,5-Me), 0.95 (d, <sup>3</sup>J<sub>H-H</sub> = 8Hz, 3H, 1-Me); <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 14.18, 20.98, 27.43, 38.04, 54.77, 61.26, 169.37 and 212.25. [Manriquez et al. *Inorg Synth XXI* 181 1982, Feilter et al. *Inorg Chem* **15** 466 1976, Threlkel et al. *Org Synth* **65** 42 1987.]

**Perfluorocyclobutane (octafluorocyclobutane)** [115-25-3] M 200.0, m -40°, b -5°, d<sup>20</sup> 1.654, d° 1.72. Purify octafluorocyclobutane by trap-to-trap distillation, retaining the middle portion. [Danus *Ind Eng Chem* **47** 144 1955, Claasen *J Chem Phys* **18** 543 1950, *Beilstein* **5** III 8, **5** IV 8.]

**Perfluorocyclohexane (dodecafluorocyclohexane)** [355-68-0] M 300.1, m 51° (sublimes), sublimes on melting at 52°, m 58.2° (sealed tube), d<sup>25</sup><sub>4</sub> 1.720, n<sup>30</sup><sub>D</sub> 1.269. Extract it repeatedly with MeOH, then pass it through a column of silica gel (previously activated by heating at 250°). [Haszeldine & Smith *J Chem Soc* 2691 1950, IR: Thompson & Temple *J Chem Soc* 1432 1948, *Beilstein* **5** III 37, **5** IV 48.]

**Perfluoro-1,3-dimethylcyclohexane** [335-27-3] M 400.1, b 101°, d<sup>20</sup><sub>4</sub> 1.829, n<sup>20</sup><sub>D</sub> 1.300. Fractionally distil it, then 35ml are sealed with about 7g KOH pellets in a borosilicate glass ampoule and heated at 135° for 48 hours. The ampoule is cooled, opened, and the liquid is resealed with fresh KOH in another ampoule and heated as before. This process is repeated until no further decomposition is observed. The substance is then washed with distilled water, dried (CaSO<sub>4</sub>) and distilled. [Grafstein *Anal Chem* **26** 523 1954, *Beilstein* **5** III 378.] **IRRITANT.**

**Perfluoro(methylcyclohexane)** [355-02-2] M 350.1, b 76.3°, d<sup>25</sup> 1.7878. Reflux it for 24 hours with saturated acid KMnO<sub>4</sub> (to oxidise and remove hydrocarbons), then neutralise, steam distil, dry with P<sub>2</sub>O<sub>5</sub> and pass slowly through a column of dry silica gel. [Glew & Reeves *J Phys Chem* **60** 615 1956.] It can also be purified by percolation through a 1 metre neutral activated alumina column, and the impurities are checked by <sup>1</sup>H NMR. [*Beilstein* **5** IV 102.] **IRRITANT.**

**R(-)- $\alpha$ -Phellandrene (*p*-menta-1,5-diene)** [4221-98-1] M 136.2, b 61°/11mm, 175-176°/760mm, d<sup>20</sup><sub>4</sub> 0.838, n<sup>20</sup><sub>D</sub> 1.471, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -230° (c 10, Et<sub>2</sub>O), -153° to -183° (neat). Purify it by gas chromatography with an Apiezon column. Also purify it by steam distillation (with 0.5% hydroquinone), then re-distil it through a 50

plate bubble cap column and collecting the fraction with **b** 72-72.5°/22mm [Pines & Eschinazi *J Am Chem Soc* 77 6318 1955]. UV:  $\lambda_{\max}$  263nm ( $\epsilon$  3,345) in octane. [Read & Storey *J Chem Soc* 2770 1930, *Beilstein* 5 III 341, 5 IV 436.]

**Picrotoxin (cocculin)** [124-87-8] **M** 602.6, **m** 203°,  $[\alpha]_{546}^{20}$   $-40^\circ$  (c 1, EtOH),  $[\alpha]_{\text{D}}^{16}$   $-29.3^\circ$  (c 4, EtOH). It is the toxic principle in the seeds of *Anamirta cocculus*. Crystallise picrotoxin from H<sub>2</sub>O or Me<sub>2</sub>CO/H<sub>2</sub>O. The *monoacetate* has **m** 244-245° (\*C<sub>6</sub>H<sub>6</sub>). [Meyer & Bruger *Chem Ber* 31 2958 1898, Johns et al. *J Chem Soc* 4717 1956, *Beilstein* 19 III/IV 5245.] HIGHLY TOXIC, stimulates the CNS and respiration.

**1R,5S- $\alpha$ -Pinene** [7785-70-8] **M** 136.2, **b** 61°/30mm, 156.2°/760mm,  $d_4^{20}$  0.858,  $n_{\text{D}}^{15}$  1.4634,  $n_{\text{D}}^{20}$  1.4658,  $[\alpha]_{\text{D}}^{20}$  +51° (neat). It is isomerised by heat, acids and certain solvents. It should be distilled under reduced pressure under N<sub>2</sub> and stored in the dark. It has been purified *via* the nitrosochloride [Waterman et al. *Recl Trav Chim, Pays Bas* 48 1191 1929]. For purification of optically active forms see Lynn [*J Am Chem Soc* 91 361 1919].

Small quantities (0.5ml) have been purified by GLC using helium as carrier gas and a column at 90° packed with 20 wt% of polypropylene sebacate on a Chromosorb support. Larger quantities are fractionally distilled under reduced pressure through a column packed with stainless steel gauze spirals. The material can be dried over CaH<sub>2</sub> or sodium, and stored in a refrigerator: CaSO<sub>4</sub> and silica gel are not satisfactory because they induce spontaneous isomerisation. [Bates et al. *J Chem Soc* 1521 1962, *Beilstein* 5 III 366, 5 IV 452.]

**1S,5S- $\alpha$ -Pinene** [7785-26-4] **M** 136.2, **b** 155-156°/760mm,  $d_4^{20}$  0.858,  $n_{\text{D}}^{20}$  1.4634,  $[\alpha]_{\text{D}}^{20}$   $-47.2^\circ$ . Purify as for 1R,5S- $\alpha$ -Pinene above. [*Beilstein* 5 III 366, 5 IV 455.]

**R(+)-Pulegone** [89-82-7] **M** 152.2, **b** 69.5°/5mm,  $d_4^{20}$  0.936,  $n_{\text{D}}^{20}$  1.4866,  $[\alpha]_{546}^{20}$  +23.5°(neat),  $[\alpha]_{\text{D}}^{20}$  +24.2°(neat). Purify pulegone *via* the *semicarbazone* which has **m** 174° (from MeOH) and  $[\alpha]_{\text{D}}^{20}$  +68.2° (c 1, CHCl<sub>3</sub>). Fractionally distil it *in vacuo*. [Short & Read *J Chem Soc* 1309 1939]. [Erskine & Waight *J Chem Soc* 3425 1960, *cf* Ort *Org Synth* 65 203 1987, *Beilstein* 7 III 334, 7 IV 188.]

**1R,3R,4R,5R-Quinic acid (1,3,4,5-tetrahydroxycyclohexane carboxylic acid)** [77-95-2] **M** 192.3, **m** 172°(dec),  $[\alpha]_{546}^{20}$   $-51^\circ$  (c 20, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{23}$   $-45^\circ$  (c 5, H<sub>2</sub>O), **pK**<sup>25</sup> 3.58. Quinic acid crystallises from H<sub>2</sub>O with **m** 174°, and from EtOH with **m** 168-169°. [McComsey & Maryanoff *J Am Chem Soc* 59 2652 1994, **pK**: Timberlake *J Chem Soc* 2795 1959, Anet & Reynolds *Aust J Chem* 8 282 1955, *Beilstein* 10 III 2407, 10 IV 2257.]

**Reductic acid (1,2-dihydroxycyclopent-1,2-en-3-one)** [80-72-8] **M** 114.1, **m** 213°, **pK**<sub>1</sub><sup>20</sup> 4.80, **pK**<sub>2</sub><sup>20</sup> 12.9. Crystallise reductic acid from EtOH, EtOAc (**m** 213-213.5°) or EtOH/EtOAc. It has been sublimed at 0.5mm. The *osazone* has **m** 245°(dec) (from BuOH). [Hess et al. *Justus Liebigs Ann Chem* 563 31 1939, 592 137 1955, 736 134 1970, *Beilstein* 8 III 1942, 8 IV 1714.]

**Squaric acid (3,4-dihydroxy-3-cyclobutene-1,2-dione)** [2892-51-5] **M** 114.1, **m** 293°(dec), 294°(dec), >300°, **pK**<sub>1</sub><sup>20</sup> 1.50, **pK**<sub>2</sub><sup>20</sup> 2.93. Purify squaric acid by recrystallisation from H<sub>2</sub>O — this is quite simple because the acid is ~7% soluble in boiling H<sub>2</sub>O and only 2% at room temperature. It is not soluble in Me<sub>2</sub>CO or Et<sub>2</sub>O; hence it can be rinsed with these solvents and dried in air or a vacuum. It is not hygroscopic and gives an intense purple colour with FeCl<sub>3</sub>. It has IR with  $\nu_{\max}$  at 1820 (C=O) and 1640 (C=C) cm<sup>-1</sup>, and UV with  $\lambda_{\max}$  at 269.5nm ( $\epsilon$  37K M<sup>-1</sup>cm<sup>-1</sup>). [Cohn et al. *J Am Chem Soc* 81 3480 1959, Park et al. *J Am Chem Soc* 84 2919 1962.] See also **pKa** values of 0.59 ± 0.09 and 3.48 ± 0.023 [Scwartz & Howard *J Phys Chem* 74 4374 1970]. [Schmidt & Reid *Synthesis* 869 1978, *Beilstein* 8 IV 2701.]

**Terpin hydrate** [2451-01-6 *cis*-hydrate, 565-50-4 and 565-48-0 *stereoisomers*] **M** 190.3, **m** 105.5° (*cis* *anhydrous*), 116-117° (*cis* *hydrate*), 156-158°, 157.5°(*trans*). Crystallise terpin from H<sub>2</sub>O or EtOH. The

anhydrous *cis*-isomer distils at 258°/760mm, but hydrates on exposure to moist air. *Anhydrous* terpin is also obtained by recrystallisation from absolute EtOH. [Sword *J Chem Soc* 127 1632 1925, Lombard & Ambrose *Bull Soc Chim Fr* 230 1961, *Beilstein* 5 IV 435.]

**1,1,2,2-Tetrafluorocyclobutane** [374-12-9] **M 128.1, b 50-50.7°, d<sub>4</sub><sup>20</sup> 1.275, n<sub>D</sub><sup>20</sup> 1.3046.** Purify 1,1,2,2-tetrafluorocyclobutane by distillation or by preparative gas chromatography using a 2m x 6mm(i.d.) column packed with β,β'-oxydipropionitrile on Chromosorb P at 33°. [Conlin & Fey *J Chem Soc, Faraday Trans 1* 76 322 1980, Coffmann et al. *J Am Chem Soc* 71 490 1949, *Beilstein* 5 III 8, 5 IV 8.]

**2,2,4,4-Tetramethylcyclobutan-1,3-dione** [933-52-8] **M 140.2, m 114.5-114.9°.** Crystallise the dione from \*C<sub>6</sub>H<sub>6</sub> and dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub> in an Abderhalden pistol. [*Beilstein* 7 III 3234, 7 IV 2004.]

**3,3,5,5-Tetramethylcyclohexanone** [14376-79-5] **M 154.3, m 11-12°, 13.2°, b 59-61°/5mm, 80-82°/13mm, 196°/760mm, 203.8-204.8°/760mm, d<sub>4</sub><sup>20</sup> 0.8954, n<sub>D</sub><sup>20</sup> 1.4515.** Purify the ketone first through a 24 inch column packed with Raschig rings, then a 40cm Vigreux column under reduced pressure (b 69-69.3°/7mm). The *oxime* has m 144-145° (from 60% EtOH), and the *semicarbazone* has m 196-197°, 197-198° (214.5°, 217-218°) [Karasch & Tawney *J Am Chem Soc* 63 2308 1941, UV: Sandris & Ourisson *Bull Soc Chim Fr* 958 1956]. [*Beilstein* 7 III 163, 7 IV 89.]

**(1R)-(-)-Thiocamphor (1R-bornane-2-thione, 1R-(-)-1,7,7-trimethylbicyclo[2.2.1]heptane-2-thione)** [53402-10-1] **M 168.3, m 136-138°, 146°, [α]<sub>D</sub><sup>22</sup> -22° (c 3, EtOAc).** It forms red prisms from EtOH and sublimes under vacuum. It possesses a sulfurous odour and is volatile like camphor. [Sen *J Indian Chem Soc* 12 647 1935, Sen *J Indian Chem Soc* 18 76 1941.] The *racemate* crystallises from \*C<sub>6</sub>H<sub>6</sub> and has m 145° [138.6-139°, White & Bishop *J Am Chem Soc* 62 10 1940]. [*Beilstein* 7 III 419.]

**1r,2t,4t-Trimethylcyclohexane** [2234-75-7] **M 126.2, b 145.7-146.7°/760mm, d<sub>4</sub><sup>20</sup> 0.786, n<sub>D</sub><sup>20</sup> 1.4330.** Wash the trimethylcyclohexane with conc H<sub>2</sub>SO<sub>4</sub> (removes aromatic hydrocarbons), then with H<sub>2</sub>O, dry it (type 4A molecular sieves), and fractionally distil it through a glass helices packed column with partial take-off and reflux ratio between 50 and 75. **Flammable liquid.** [cf. Henne et al. *J Am Chem Soc* 63 3475 1941, Rossini *Anal Chem* 20 112 1948, *Beilstein* 5 H 42, 5 I 17, 5 II 24, 5 III 121.]

**R-(-)-2,2,6-Trimethyl-1,4-cyclohexanedione** [60046-49-3] **M 154.2, m 88-90°, 91-92°, [α]<sub>D</sub><sup>20</sup> -270° (c 0.4%, MeOH), [α]<sub>D</sub><sup>20</sup> -275° (c 1, CHCl<sub>3</sub>).** It is obtained from fermentation and is purified by recrystallisation from diisopropyl ether. [ORD: Leuenberger et al. *Helv Chim Acta* 59 1832 1976.] The *racemate* has m 65-67°, and the 4-(4-phenyl)semicarbazone has m 218-220° (from CH<sub>2</sub>Cl<sub>2</sub>/MeOH) [Isler et al. *Helv Chim Acta* 39 2041 1956, *Beilstein* 7 IV 2032.]

***cis,cis*-1α,3α,5α-Trimethylcyclohexane-1,3,5-tricarboxylic acid (Kemp's acid)** [79410-20-1] **M 258.3, m 241-243°, pK<sub>1</sub> 3.30, pK<sub>2</sub> 5.85, pK<sub>3</sub> 7.3 (H<sub>2</sub>O); pK<sub>1</sub> 4.7, pK<sub>2</sub> 7.6, pK<sub>3</sub> 8.8 (50% H<sub>2</sub>O/MeOH).** Recrystallise the tricarboxylic acid from Me<sub>2</sub>CO after re-precipitating it several times with mineral acid from aqueous alkaline solution. The *trimethyl ester* has m 78-81°. [See Kemp *J Org Chem* 46 5140 1981, Jeong et al. *J Am Chem Soc* 113 201 1991, Stack et al. *J Am Chem Soc* 114 7007 1992.]

**(±)-2,2,6-Trimethylcyclohexanone** [2408-37-9] **M 140.2, b 69-71.5°/20mm, 177-178.5°/758mm d<sub>4</sub><sup>20</sup> 0.904, n<sub>D</sub><sup>20</sup> 1.4470.** Purify it *via* the *semicarbazone* (m 218°, from MeOH or EtOH), decompose this in the usual way (cf p 75 and 77, or MEK, p 123) and fractionally distil the liquid ketone through a Vigreux column at ~760mm. [Chakravarti *J Chem Soc* 1567 1947, Milas et al. *J Am Chem Soc* 70 1831 1948, *Beilstein* 7 I 24, 7 II 32, 7 III 7.]

**Xanthatin (3-methylene-7-methyl-6-[3-oxo-1-buten-1-yl]cyclohept-5-ene-[10,11-*b*]furan-2-one, (-)-2-[(1R)-7*t*-hydroxy-5*c*-methyl-4-(3-oxobut-1-en-ξ-yl)cyclohept-3-en-*r*-yl]-acrylic acid lactone** [26791-73-1] **M 246.3, m 114.5-115°, [α]<sub>D</sub><sup>20</sup> -20° (c 2, CHCl<sub>3</sub>).** Crystallise xanthatin from MeOH, aqueous MeOH, EtOH

or aqueous EtOH. Its UV has  $\lambda_{\max}(\epsilon)$  at 213 (22800) and 275nm (7300). The 2,4-dinitrophenylhydrazone has **m** 240°(dec) (twice recrystallise from  $\text{CHCl}_3/\text{MeOH}$ ). [Geissman et al. *J Am Chem Soc* **76** 685 1954, Deuel & Geissman *J Am Chem Soc* **79** 3778 1957, *Beilstein* **17** III/IV 6221, **17/1** V 305.]

## AROMATIC COMPOUNDS

**Acenaphthene** [83-32-9] **M 154.2, m 94.0°**. Crystallise acenaphthene from EtOH. It has also been purified by chromatography from CCl<sub>4</sub> on alumina with \*benzene as eluent [McLaughlin & Zainal *J Chem Soc* 2485 1960]. [*Beilstein* 5 IV 1834.]

**Acenaphthenequinone** [82-86-0] **M 182.2, m 260-261°**. Extract it with, then recrystallise it twice from \*C<sub>6</sub>H<sub>6</sub>. Dry it *in vacuo*. [LeFevre et al. *J Chem Soc* 974 1963, *Beilstein* 7 IV 2498.]

**RS-Acenaphthenol** [6306-07-6] **M 170.2, m 144.5-145.5°, 146°, 148°**. If highly coloured (yellow), dissolve it in boiling \*benzene (14g in 200ml), add charcoal (0.5g), filter it through a heated funnel, concentrate to 100ml and cool to give almost colourless needles. \*Benzene vapour is **TOXIC**; use an efficient fume cupboard. The acetate has **m** 166-168°/5mm (bath temperature 180-185°). [Cason *Org Synth Coll Vol III* 3 1955.] It can also be recrystallised from \*C<sub>6</sub>H<sub>6</sub> or EtOH [Fieser & Cason *J Am Chem Soc* 62 432 1940]. It forms a brick-red crystalline **complex** with 2,4,5,7-tetranitrofluoren-9-one which is recrystallised from AcOH and is dried in a vacuum over KOH and P<sub>2</sub>O<sub>5</sub> at room temperature, **m** 170-172° [Newman & Lutz *J Am Chem Soc* 78 2469 1956]. [*Beilstein* 6 IV 4623.]

**Acenaphthylene** [208-96-8] **M 152.2, m 92-93°, b 280°/~760mm**. Dissolve acenaphthylene in warm redistilled MeOH, filter through a sintered glass funnel and cool to -78° to precipitate the material as yellow plates [Dainton et al. *Trans Faraday Soc* 56 1784 1960]. Alternatively it can be sublimed *in vacuo*. [*Beilstein* 5 H 625, 5 IV 2138.]

**4-Acetamidobenzaldehyde** [122-85-0] **M 163.2, m 155°, 156°, 160°**. Recrystallise it from water. The 4-nitrophenylhydrazone, **m** 264-265°, crystallises as orange needles from EtOH [Hodgson & Beard *J Chem Soc* 21 1927, *Beilstein* 14 H 38, 14 II 25, 14 III 75, 14 IV 71.]

**p-Acetamidobenzenesulfonyl chloride (N-acetylsulfanilyl chloride)** [121-60-8] **M 233.7, m 149°(dec)**. Crystallise the chloride from toluene, CHCl<sub>3</sub>, or ethylene dichloride. [*Beilstein* 14 IV 2703.]

**α-Acetamidocinnamic acid** [5469-45-4] **M 205.2, m 185-186° (2H<sub>2</sub>O), 190-191°(anhydrous), 193-195°, pK<sub>Est</sub> ~3.2**. It crystallises from H<sub>2</sub>O as the *dihydrate*, and on drying at 100° it forms the *anhydrous* compound which is *hygroscopic*. Alkaline hydrolysis yields NH<sub>3</sub> and phenylpyruvic acid. [Erlenmeyer & Früstück *Justus Liebigs Ann Chem* 284 47 1895, *Beilstein* 14 IV 1769.]

**2-Acetamidofluorene (N-[2-fluorenyl]acetamide)** [53-96-3] **M 223.3, m 194°, 196-198°**. Recrystallise it from toluene (1.3mg in 100ml). Its solubility in H<sub>2</sub>O is 1.3mg/l at 25°, its UV has λ<sub>max</sub> at nm(log ε): 288(4.43), 313(4.13). [Sawicki *J Org Chem* 21 271 1956.] It can also be recrystallised from 50% AcOH. [Diels et al. *Chem Ber* 35 3285 1902]. 9-<sup>14</sup>C and ω-<sup>14</sup>C 2-acetamidofluorene were recrystallised from aqueous EtOH and had **m** 194-195° and 194° respectively. **Potent CARCINOGEN**. [Miller et al. *Cancer Res* 9 504 1949, 10 616 1950, Sadin et al. *J Am Chem Soc* 74 5073 1952, *Beilstein* 12 H 3287, 12 IV 3373.]

**2-Acetamidophenol** [614-80-2] **M 151.2, m. 209°, pK<sub>Est</sub> ~9.4**. Recrystallise it from water, EtOH or aqueous EtOH. [*Beilstein* 13 H 370, 13 I 113, 13 II 171, 13 III 778.]

**3-Acetamidophenol (Metacetamol)** [621-42-1] **M 151.2, m 148-149°, pK<sup>25</sup> ~9.59**. Recrystallise the phenol from water. The 3,5-dinitrobenzamide **complex** gives orange-yellow crystals from hot H<sub>2</sub>O and has **m** 212°. [*Beilstein* 13 H 415, 13 I 132, 13 II 213, 13 III 950, 13 IV 977.]

**4-Acetamidophenol (Paracetamol, acetaminophen, 4'-hydroxyacetanilide)** [103-90-2] **M 151.2, m 169-170.5°, pK<sub>Est</sub> ~10.0**. Recrystallise Paracetamol from water or EtOH. The 3,5-dinitrobenzamide **complex** gives orange crystals from hot H<sub>2</sub>O and has **m** 171.5°. [*Beilstein* 13 H 460, 13 I 159, 13 II 243, 13 III 1056, 13 IV

1091.]

***p*-Acetamidophenylacetic acid (Actarit)** [18699-02-0] **M 193.2, m 167°, 168-170°, 174-175°, pK<sup>25</sup> 3.49.** Crystallise the acid from MeOH/Me<sub>2</sub>CO, aqueous EtOH or H<sub>2</sub>O. The *amide* has **m 231°** (from 50% aqueous EtOH). [Gabriel *Chem Ber* **15** 841 1882, Cerecedo et al. *J Biol Chem* **42** 238 1924, Tramontano et al. *J Am Chem Soc* **110** 2282 1988, *Beilstein* **14** II 281.]

**Acetanilide** [103-84-4] **M 135.2, m 114°, pK<sup>25</sup> 0.5.** Recrystallise acetanilide from water, aqueous EtOH, \*benzene or toluene. [*Beilstein* **12** IV 373.]

**Acetoacetanilide** [102-01-2] **M 177.2, m 86°, pK<sup>25</sup> 10.68.** Crystallise the anilide from H<sub>2</sub>O, aqueous EtOH or petroleum ether (b 60-80°). [Williams & Krynitsky *Org Synth Coll Vol III* 10 1955.]

**4-Acetophenetidide (phenacetin, *p*-methoxyacetanilide)** [62-44-2] **M 179.2, m 136°.** Crystallise it from H<sub>2</sub>O or EtOH, and its solubility in H<sub>2</sub>O is 0.08% (at ~10°) and 1.2% (at ~100°), and in EtOH it is 6.7% (at ~10°) and 36% (at ~100°). *Alternatively*, it can be purified by solution in cold dilute alkali and re-precipitating by addition of acid to neutralisation point. Dry it in air. [*Beilstein* **13** H 461, **13** IV 1092.]

**Acetophenone** [98-86-2] **M 120.2, m 19.6°, b 54°/2.5mm, 202°/760mm, d<sub>4</sub><sup>25</sup> 1.0238, n<sub>D</sub><sup>25</sup> 1.5322, pK<sup>26</sup> -7.6(basic), pK<sup>25</sup> 19.2(acidic).** Dry it by fractional distillation or by standing with anhydrous CaSO<sub>4</sub> or CaCl<sub>2</sub> for several days, followed by fractional distillation under reduced pressure (from P<sub>2</sub>O<sub>5</sub>, optional), and careful, slow and repeated partial crystallisations from the liquid at 0° excluding light and moisture. It can also be crystallised at low temperatures from isopentane. Distillation can be followed by purification using gas-liquid chromatography [Earls & Jones *J Chem Soc, Faraday Trans 1* **71** 2186 1975.] [*Beilstein* **7** H 271, **7** I 146, **7** II 208, **7** III 936, **7** IV 619.]

§ A commercial polystyrene supported version is available — scavenger resin (for diol substrates).

**Aceto-*o*-toluidide (2-methylacetanilide)** [120-66-1] **M 149.2, m 110°, 112°, b 176°/14mm, 296°/760mm.** Crystallise the toluidide from hot H<sub>2</sub>O (solubility 1g/210ml), EtOH or aqueous EtOH. Its UV has  $\lambda_{\max}$  at 230 and 280nm (EtOH). [*Beilstein* **12** H 792, **12** I 376, **12** II 439, **12** III 1853, **12** IV 1755.]

**Aceto-*m*-toluidide (3-methylacetanilide)** [537-92-8] **M 149.2, m 65.5°, b 182-183°/14mm, 303°/760mm.** Crystallise the toluidide from H<sub>2</sub>O, EtOH, aqueous EtOH or Et<sub>2</sub>O/petroleum ether (**m 66°**). Its UV has  $\lambda_{\max}$  at 245nm (EtOH). [*Beilstein* **12** H 860, **12** I 400, **12** II 468, **12** III 1962, **12** IV 1823.]

**Aceto-*p*-toluidide (4-methylacetanilide)** [103-89-9] **M 149.2, m 146°, b 307°/760mm.** Crystallise it from aqueous EtOH. [*Beilstein* **12** H 920, **12** I 420, **12** II 501, **12** III 2051, **12** IV 1902.]

***R*-(-)- $\alpha$ -Acetoxyphenylacetic (acetyl mandelic) acid** [51019-43-3] **M 194.2, m 96-98°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -153.7° (c 2.06, Me<sub>2</sub>CO), [ $\alpha$ ]<sub>546</sub><sup>20</sup> -194° (c 2.4, Me<sub>2</sub>CO), pK<sub>Est</sub> ~2.9.** It crystallises from H<sub>2</sub>O with 1mol of solvent which is removed on drying, or from other solvents as for the *S*-isomer below. [Angus & Owen *J Chem Soc* 227 1943, Parker *Chem Rev* **91** 1441 1991, *Beilstein* **10** III 453.]

***S*-(+)- $\alpha$ -Acetoxyphenylacetic (acetyl mandelic) acid** [7322-88-5] **M 194.2, m 80-81°, 95-97.5°, [ $\alpha$ ]<sub>D</sub><sup>27</sup> +158° (c 1.78, Me<sub>2</sub>CO), [ $\alpha$ ]<sub>546</sub><sup>20</sup> +186° (c 2, Me<sub>2</sub>CO).** Recrystallise it from \*benzene/hexane or toluene, and it has characteristic NMR and IR spectra. [Pracejus *Justus Liebigs Ann Chem* **622** 10 1959, Breitholle & Stammer *J Org Chem* **39** 1311 1974, *Beilstein* **10** IV 567.]

**9-Acetylanthracene** [784-04-3] **M 220.3, m 75-76°.** Crystallise 9-acetylanthracene from EtOH. [Masnori et al. *J Am Chem Soc* **108** 1126 1986, *Beilstein* **7** II 450.]

***N*-Acetylanthranilic acid** [89-52-1] **M 179.1, m 182-184°, 185-186°, 190°(dec), pK<sup>20</sup> 3.61.** Wash the acid with distilled H<sub>2</sub>O and recrystallise it from aqueous AcOH, dry it and recrystallise again from EtOAc. Also recrystallise it from water or EtOH. Its UV has  $\lambda_{\max}$  at 221, 252 and 305nm (EtOH). The *amide* crystallises

from aqueous EtOH and has **m** 186-187° and  $\lambda_{\max}$  218, 252 and 301nm. [Chattaway *J Chem Soc* 2495 1931, Walker *J Am Chem Soc* 77 6698 1955, *Beilstein* 14 H 337, 14 I 540, 14 II 219, 14 III 922.]

**2-Acetylbenzoic acid** [577-56-0] **M 164.2, m 115-116°, 116-118°, pK<sup>20</sup> 4.14, pK<sup>25</sup> 4.10.** It crystallises from \*C<sub>6</sub>H<sub>6</sub> and H<sub>2</sub>O (15g/100ml). The *oxime* has **m** 156-157°, and the *2,4-dinitrophenylhydrazone* has **m** 185-186°(needles from EtOH). [Yale *J Am Chem Soc* 69 1547 1947, Panetta & Miller *Synthesis* 43 1977, *Beilstein* 10 H 690, 10 I 330, 10 II 479, 10 III 3025, 10 IV 2766.]

**4-Acetylbenzoic acid** [586-89-0] **M 164.2, m 207.5-209.5°, 208.6-209.4°, pK<sup>25</sup> 3.70, 5.21, 5.10 (EtOH).** Dissolve the acid in 5% aqueous NaOH, extract it with Et<sub>2</sub>O, and acidify the aqueous solution. Collect the precipitate, and recrystallise it from boiling H<sub>2</sub>O (100 parts) using decolorising charcoal [Pearson et al. *J Org Chem* 24 504 1959, Pearson et al. *J Chem Soc* 265 1957, Detweiler & Amstutz *J Am Chem Soc* 72 2882 1950, Bordwell & Cooper *J Am Chem Soc* 74 1058 1952]. [*Beilstein* 10 IV 2769.]

**4-Acetylbenzotrile** [1443-80-7] **M 145.2, m 57-58°.** Recrystallise the nitrile from EtOH [Wagner et al. *J Am Chem Soc* 108 7727 1986]. [*Beilstein* 10 H 695, 10 III 3030.]

**Acetyl-5-bromosalicylic acid** [1503-53-3] **M 259.1, m (156°), 168°, 168-169°, pK<sub>Est</sub> ~3.0.** Crystallise the acid from EtOH. [Robertson *J Chem Soc* 81 1482 1902, *Beilstein* 10 H 108, 10 II 64.]

**2-Acetylfluorene** [781-73-7] **M 208.3, m 130-131°, 132°.** Crystallise acetylfluorene from EtOH (solubility is 60g/800ml) or Me<sub>2</sub>CO (solubility is 60g/400ml). The *oxime* [110827-07-1] has **m** 192-193.5° and the *2,4-dinitrophenylhydrazone* [109682-26-0] has **m** 261-262°. [Ray & Rieveschl *Org Synth Coll Vol III* 23 1973.]

**5(3)-Acetyl-2(6)-methoxybenzaldehyde** [531-99-7] **M 166.2, m 144°.** Extract a solution of the aldehyde in \*C<sub>6</sub>H<sub>6</sub> with 20% aqueous sodium bisulfite, and the bisulfite adduct in the aqueous solution is decomposed by acidifying and heating whereby the aldehyde separates. It is collected, washed with H<sub>2</sub>O, dried in a vacuum. It is recrystallised from EtOH (**m** 140-141°) and then from Et<sub>2</sub>O (**m** 143-144°). [Gray & Bonner *J Am Chem Soc* 70 1249 1948, Angyal et al. *J Chem Soc* 2142 1950, *Beilstein* 8 IV 1984.]

**4-Acetyl-N-methylaniline ([4-methylamino]acetophenone)** [17687-47-7] **M 149.2, m 102-106°, 103-107°.** This herbicide crystallises from H<sub>2</sub>O. The *4-acetyl-N,N-dimethylaniline* derivative forms colourless plates also from H<sub>2</sub>O with **m** 58-59°. [Klingel *Chem Ber* 18 2694 1885, Staudinger & Kon *Justus Liebigs Ann Chem* 384 111 1911, *Beilstein* 14 H 47, 14 I 366.]

**1-Acetylnaphthalene (1-acetonaphthenone)** [941-98-0] **M 170.1, m 10.5°, b 93-95°/0.1mm, 167°/12mm, 302°/760mm, d<sub>4</sub><sup>20</sup> 1.12, pK<sup>25</sup> -6.22 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>).** If the NMR spectrum indicates the presence of impurities, probably 2-acetylnaphthalene, convert the substance to its *picrate* by dissolving in \*benzene or EtOH and adding excess of saturated picric acid in these solvents until separation of picrates is complete. Recrystallise the picrate till the melting point is 118°. Decompose the picrate with dilute NaOH and extract with Et<sub>2</sub>O. Dry the extract (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and distil the residue. The *2,4-dinitrophenylhydrazone* crystallises from EtOH and has **m** 259°. [Stobbe & Lenzer *Justus Liebigs Ann Chem* 380 95 1911, Williams & Osborne *J Am Chem Soc* 61 3438 1939, *Beilstein* 7 IV 1292.]

**2-Acetylnaphthalene (2-acetonaphthenone, β-acetonaphthone, 2-acetonaphthalene, methyl-2-naphthylketone)** [93-08-3] **M 170.2, m 52-53°, 55°, 55.8°, b 164-166°/8mm, 171-173°/17mm, 301-303°/760mm, pK<sup>25</sup> -6.16 (H<sub>0</sub> scale, aqueous H<sub>2</sub>SO<sub>4</sub>).** Separate it from the 1-isomer by fractional crystallisation of the *picrate* in EtOH (see entry for the 1-isomer above) to **m** 82°. Decomposition of the picrate with dilute NaOH and extraction with Et<sub>2</sub>O, then evaporation, give purer 2-acetylnaphthalene. If this residue solidifies, it can be recrystallised from petroleum ether, EtOH or acetic acid; otherwise it should be distilled in a vacuum and the solid distillate is recrystallised [Gorman & Rodgers *J Am Chem Soc* 108 5074 1986, Levanon et al. *J Phys Chem* 91 14 1987]. Purity should be checked by high field NMR spectroscopy. Its *oxime* has **m** 145°(dec), and the *semicarbazone* has **m** 235°. [Stobbe & Lenzer *Justus Liebigs Ann Chem* 380 95 1911, Raffauf *J Am Chem Soc* 72 753 1950, Hunsberger *J Am Chem Soc* 72 5626 1950, Immediata & Day *J Org Chem* 5 512 1940,

*Beilstein* 7 IV 1294.]

**1-Acetyl-2-phenylhydrazine** [114-83-0] **M 150.2, m 128.5°, pK<sup>25</sup> 1.3.** Crystallise the hydrazine from aqueous EtOH. [*Beilstein* 15 H 241.]

**Acetylsalicylic acid (Aspirin)** [50-78-2] **M 180.2, m 133.5-135° (and various), pK<sup>25</sup> 3.38, (pK<sup>17</sup> 3.56).** Crystallise aspirin twice from toluene, wash it with cyclohexane and dry it at 60° under vacuum for several hours [Davis & Hetzer *J Res Nat Bur Stand* 60 569 1958]. It has been recrystallised from isopropanol and from diethyl ether/petroleum ether (b 40-60°). It crystallises from EtOH (m 143-144°), \*C<sub>6</sub>H<sub>6</sub> (m 143°), hexane (m 115° and 128°), octane (m 121°), and has m 110° after sublimation. It has pK<sup>26</sup> 3.69(H<sub>2</sub>O), 4.15(20% aqueous EtOH), 4.47(30% aqueous EtOH) and 4.94(40% aqueous EtOH). It is an analgesic. [*Beilstein* 10 H 67, 10 II 41, 10 III 102, 10 IV 138.]

**O-Acetylsalicyloyl chloride** [5538-51-2] **M 198.6, m 45°, 46-49°, 48-52°, b 107-110°/0.1mm, 115°/5mm, 135°/12mm, n<sub>D</sub><sup>20</sup> 1.536.** Check first the IR to see if an OH frequency is present. If so, some free acid is present. Then reflux with acetyl chloride for 2-3 hours and fractionate at high vacuum. The distillate should crystallise. It can be recrystallised from hexane or \*C<sub>6</sub>H<sub>6</sub> (m 60°, sintering at 52°). [Riegel & Wittcoff *J Am Chem Soc* 64 486 1942, *Beilstein* 10 H 86, 10 I 43, 10 II 55, 10 III 151, 10 IV 169.]

**O-Acetylsalicylsalicylic acid (Salsalate acetate)** [530-75-6] **M 300.3, m 159°.** Crystallise the analgesic from dilute AcOH or EtOH (m 161-162°), MeOH (m 165-168°), and \*C<sub>6</sub>H<sub>6</sub>/EtOH (m 163-165°). Its solubilities in boiling Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> and EtOH are 1.4%, 2.2% and 33%, respectively. [Baker et al. *J Chem Soc* 201 1951, Garrett et al. *J Am Pharm Soc* 48 684 1959, *Beilstein* 10 I 41, 10 II 54, 10 IV 165.]

**N-(4-Acetylsulfanilamide (sulfacetamide)** [144-80-9] **M 214.2, m 216°.** Crystallise the amide from aqueous EtOH. [*Beilstein* 14 IV 2662.]

**Acetyl *p*-toluenesulfonate** [26908-82-7] **M 214.2, m 54-56°, b 186-188°/20mm.** The most likely impurity is *p*-toluenesulfonic acid (could be up to 10%). This can be removed by dissolving it in dry Et<sub>2</sub>O and cooling until the anhydride crystallises out. It decomposes on heating; below ~130° it gives the disulfonic anhydride and above ~130° polymers are formed, but it can be distilled in a vacuum if it is free of acid. It is used for cleaving ethers [Prep, IR, NMR: Karger & Mazur *J Org Chem* 36 528, Karger & Mazur *J Org Chem* 36 532 1971]. [*Beilstein* 11 III 255.]

**Allyl phenyl sulfide** [5296-64-0] **M 150.2, b 59-60°/1.5mm, 79-80°/3mm, 114-114.3°/23.5mm, 225-226°/740mm, 215-218°/750mm, d<sub>4</sub><sup>20</sup> 1.0275, n<sub>D</sub><sup>20</sup> 1.5760.** Dissolve the sulfide in Et<sub>2</sub>O, wash with alkali, H<sub>2</sub>O, dry over CaCl<sub>2</sub>, evaporate and fractionally distil it, preferably under vacuum. It should not give a precipitate with an alcoholic solution of Pb(OAc)<sub>2</sub>. [Hurd & Greengard *J Am Chem Soc* 52 3356 1930, Tarbell & McCall *J Am Chem Soc* 74 48 1952, *Beilstein* 6 IV 1479.]

**Amberlite IRA-904 Anion exchange resin (Rohm and Haas)** [9050-98-0]. Wash with 1M HCl, CH<sub>3</sub>OH (1:10) and then rinse it with distilled water until the washings are neutral to litmus paper. Finally extract successively for 24 hours in a Soxhlet apparatus with MeOH, \*benzene and cyclohexane [Shue & Yan *Anal Chem* 53 2081 1981]. It is a strong basic resin also used for base catalysis [Fieser & Fieser *Reagents for Org Synth* 1 511, Wiley 1967].

***p*-Aminoacetanilide** [122-80-5] **M 150.2, m 162-163°, 163°, 165-167°, 166-167°, pK<sup>15</sup> 4.46, pK<sup>40</sup> 3.94.** Crystallise the anilide from water. It has an unstable crystalline form with m 141°. It has IR has ν<sub>max</sub> (CCl<sub>4</sub>) at 1681cm<sup>-1</sup>. [*Beilstein* 13 H 94, 13 I 28, 13 II 50, 13 III 166, 13 IV 137.]

**ω-Aminoacetophenone hydrochloride (phenacylamine hydrochloride, 2-aminoacetophenone HCl)** [5468-37-1] **M 171.6, m 188°(dec), 194°(dec), pK<sup>25</sup> 5.34.** Crystallise the salt from Me<sub>2</sub>CO /EtOH, EtOH/ Et<sub>2</sub>O, 2-propanol or 2-propanol and a little HCl (slowly after a few days). The *oxime* of the free base has m 140°, and

the *picrate* of the free base has  $m$  182° (from EtOH). [Castro *J Am Chem Soc* **108** 4179 1986, Baumgarten & Petersen *Org Synth Coll Vol V* 909 1973, cf *Beilstein* **14** H 49, **14** III 105.]

***m*-Aminoacetophenone** [99-03-6] **M 135.2, m 98-99°, b 189-290°/760mm, pK<sup>25</sup> 3.56.** Recrystallise it from EtOH or aqueous EtOH (**m** 99.5°). The *thiosemicarbazone* has **m** 202-204° (from EtOH). [*Beilstein* **14** H 45, **14** IV 96.]

***p*-Aminoacetophenone** [99-92-3] **M 135.2, m 104-106°, 105-107°, b 293°/atm, pK<sup>25</sup> 2.19.** Recrystallise it from CHCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O. It is soluble in hot H<sub>2</sub>O. Its UV (EtOH) has  $\lambda_{\max}$  at 403nm (log  $\epsilon$  4.42) [Johnson *J Am Chem Soc* **75** 2720 1953]. [Vandenbelt *Anal Chem* **26** 726 1954.] The *2,4-dinitrophenylhydrazone* has **m** 266-267° (from CHCl<sub>3</sub> or EtOH) with  $\lambda_{\max}$  at 403nm (log $\epsilon$  4.42), and the *semicarbazone* has **m** 193-194°(dec)(from MeOH). The *hydrochloride* has **m** 98°(dec)(from H<sub>2</sub>O). [*Beilstein* **14** IV 100.]

**1-Aminoanthraquinone-2-carboxylic acid** [82-24-6] **M 276.2, m 295-296°.** Crystallise the acid from nitrobenzene. It is used for the detection of Al, Mg Cd, Zn, Mn, Cu, Hg, Fe, Co, Ni and Pb. The *methyl ester* gives red needles from AcOH, **m** 228°. The *ethyl ester*, **m** 198°, crystallises also as red needles from AcOH. [Locher & Fietz *Helv Chim Acta* **10** 667 1927, *Beilstein* **14** II 419, **14** III 168.]

***p*-Aminoazobenzene (*p*-phenylazoaniline)** [60-09-3] **M 197.2, CI 11000, m 126°, pK<sup>25</sup> ~2.82.** Crystallise this dye from EtOH, CCl<sub>4</sub>, petroleum ether/\*C<sub>6</sub>H<sub>6</sub>, or a MeOH/H<sub>2</sub>O mixture. [*Beilstein* **16** IV 445.]

***o*-Aminoazotoluene (Fast Garnet GBC base, 4'-amino-2,3'dimethylazobenzene, Solvent yellow 3)** [97-56-3] **M 225.3, m 101.4-102.6°, CI 11160, pK<sup>26</sup> 2.29 (50% aqueous EtOH).** Recrystallise the dye twice from EtOH, once from \*benzene, then dry it in an Abderhalden drying apparatus. [Cilento *J Am Chem Soc* **74** 968 1952, Sawicki *J Org Chem* **21** 605 1956, *Beilstein* **16** H 334, **16** I 322, **16** II 178, **16** III 386, **16** IV 525.] **CARCINOGENIC.**

**2-Aminobenzaldehyde** [529-23-7] **M 121.1, m 39-40°, 80-82°/2mm, pK<sup>20</sup> 1.36.** Distil it in steam and recrystallise it from H<sub>2</sub>O or EtOH/Et<sub>2</sub>O. The *semicarbazone* has **m** 247°. [*Beilstein* **14** H 21, **14** I 356, **14** II 14, **14** III 47, **14** IV 42.]

**2-Aminobenzaldehyde phenylhydrazone (Nitrin)** [63363-93-9] **M 211.3, m 227-229°.** Crystallise it from acetone. [Knöpfer *Monatsh Chem* **31** 97 1910, *Beilstein* **14** H 21, **14** II 14, **14** III 47.]

**3-Aminobenzaldehyde** [29159-23-7] **M 121.1, m 28-30°, pK<sub>Est</sub> ~2.0.** The aldehyde crystallises as light yellow plates from ethyl acetate. The UV has  $\lambda_{\max}$  at 227 and 327.5nm in cyclohexane. The *acetyl* derivative has **m** 122° (from EtOH) and the *oxime* has **m** 195° (yellow-brown plates from EtOH). [*Beilstein* **14** H 28, **14** I 359, **14** II 21, **14** III 53, **14** IV 46.]

**4-Aminobenzamide hydrochloride** [59855-11-7] **M 199.6, m 284-285°, pK<sub>Est</sub> ~1.7.** Recrystallise the salt from EtOH. The *free base* [2835-68-9] **M 136.2, has m 182.9° and crystallises with 0.25H<sub>2</sub>O (m 178-179°).** [Rupe & Vogler *Helv Chim Acta* **8** 835 1925, *Beilstein* **14** H 425, **14** III 1061.]

***p*-Aminobenzeneazodimethylaniline** [539-17-3] **M 240.3, m 182-183°.** Crystallise the azo-dye from aqueous EtOH. [*Beilstein* **14** IV 1004.]

***o*-Aminobenzoic acid (anthranilic acid)** [118-92-3] **M 137.1, m 145°, pK<sub>1</sub><sup>25</sup> 2.94, pK<sub>2</sub><sup>25</sup> 4.72.** Crystallise anthranilic acid from water (charcoal). It has also been recrystallised from 50% aqueous acetic acid. It sublimes in a vacuum. [*Beilstein* **14** IV 1004.]

***m*-Aminobenzoic acid** [99-05-8] **M 137.1, m 174°, pK<sub>1</sub><sup>25</sup> 3.29, pK<sub>2</sub><sup>25</sup> 5.10.** Crystallise the acid from water. [*Beilstein* **14** IV 1092.]

***p*-Aminobenzoic acid** [150-13-0] **M 137.1, m 187-188°, pK<sub>1</sub><sup>25</sup> 2.45, pK<sub>2</sub><sup>25</sup> 4.85.** Purify *p*-aminobenzoic acid

by dissolving it in 4-5% aqueous HCl at 50-60°, decolorising with charcoal and carefully precipitating it with 30% Na<sub>2</sub>CO<sub>3</sub> to pH 3.5-4 in the presence of ascorbic acid. It can be recrystallised from water, EtOH or EtOH/water mixtures. [Beilstein 14 IV 1126.]

**p-Aminobenzonitrile (p-cyanoaniline)** [873-74-5] M 118.1, m 86-86.5°, 85-87°, pK<sup>25</sup> 1.74. It crystallises from water, 5% aqueous EtOH or EtOH and is dried over P<sub>2</sub>O<sub>5</sub>, or dried *in vacuo* for 6 hours at 40°. [Moore et al. *J Am Chem Soc* 108 2257 1986, Edidin et al. *J Am Chem Soc* 109 3945 1987, Beilstein 14 IV 1158.]

**4-Aminobenzophenone** [1137-41-3] M 197.2, m 123-124°, pK<sup>25</sup> 2.17. Dissolve it in aqueous acetic acid, filter and precipitate it with ammonia. This process is repeated several times, then the amine is recrystallised from aqueous EtOH. [Beilstein 14 IV 248.]

**2-Aminobiphenyl (o-aminobiphenyl)** [90-41-5] M 169.2, m 47-48°, 49°, 49.3°, 47-50°, b 114°/2mm, 135°/5.5mm, 160°/11mm, 166-168°/16mm, 182°/30mm, 299.0°/atm, pK<sup>18</sup> 3.85, pK<sup>20</sup> 3.39, pK<sup>20</sup> 3.03, 3.34 (50% aqueous EtOH). It is prepared by reduction of 2-nitrobiphenyl [see 86-00-0] with Fe/HCl [Morgan and Walls *J Soc Chem Ind* 49 15T 1930], or catalytically in batches of 150g by H<sub>2</sub> at 75 atmospheres pressure in the presence of Raney Ni, which begins at ~80° and is not allowed to rise above 100°, in 90% yield [Cookson & Mann *J Chem Soc* 2891 1949]. It crystallises from aqueous EtOH (charcoal). The *picrate* has m 164-165° (from EtOH or H<sub>2</sub>O), the *N-acetyl derivative* has m 118° (from EtOH), the *N-phenylsulfonate derivative* has m 292° (from MeOH), and the *N-tosylate* has m 194.1-195.6° (from MeOH). [Beilstein 12 H 1317, 12 I 546, 12 II 747, 12 III 3124, 12 IV 3223.]

**p-Aminobiphenyl** [92-67-1] M 169.2, m 53°, b 191°/16mm, pK<sup>18</sup> 4.38. Crystallise it from water or EtOH. [Beilstein 12 IV 3241.] **CARCINOGENIC.**

**5-Amino-2-bromobenzoic acid** [2840-02-0] M 216.0, m 178°, 180°, pK<sub>Est(1)</sub> ~1.7, pK<sub>Est(2)</sub> ~4.4. Crystallise the acid from H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub> (m 128°). The *acetyl derivative* crystallises from H<sub>2</sub>O (as *monohydrate*) or absolute EtOH with m 196-197° (*anhydrous*). [Koopal *Rec Trav Chim Pays Bas* 34 148 1915, Bamberger *Chem Ber* 57 2090 1924, Beilstein 14 H 413, 14 II 245.]

**2-Amino-5-bromotoluene (4-bromo-2-methylaniline)** [583-75-5] M 186.1, m 59°, 59.5°, 240°/760mm, pK<sup>25</sup> 3.58. Steam distil the aniline and recrystallise it from EtOH. It has UV with λ<sub>max</sub> at 292.5nm (H<sub>2</sub>O). [Beilstein 12 H 838, 12 I 389, 12 II 456, 12 IV 1804.]

**2-Amino-5-chlorobenzoic acid** [635-21-1] M 171.6, m 100°, pK<sub>1</sub><sup>25</sup> 1.69, pK<sub>2</sub><sup>25</sup> 4.35. Crystallise the acid from water, EtOH or chloroform. [Beilstein 14 IV 1075.]

**3-Amino-4-chlorobenzoic acid** [2840-28-0] M 171.6, m 216-217°, pK<sub>Est(1)</sub> ~2.7, pK<sub>Est(2)</sub> ~2.9. Crystallise the acid from water. [Beilstein 14 IV 1115.]

**4-Amino-4'-chlorobiphenyl** [135-68-2] M 203.5, m 132-133°, 134°, pK<sub>Est</sub> ~4.0. Crystallise the amine from petroleum ether, EtOH or aqueous EtOH. The *acetyl derivative* has m 245° from EtOH. [Dewar & James *J Chem Soc* 4270 1958, Gelmo *Chem Ber* 39 4176 1906, Beilstein 12 H 1319, 12 II 757, 12 IV 3269.]

**2-Amino-4,6-dichlorophenol** [527-62-8] M 175.0, m 95-96°, pK<sub>Est(1)</sub> ~3.1, pK<sub>Est(2)</sub> ~6.8. Crystallise the phenol from CS<sub>2</sub> or \*benzene. It sublimes at 0.06mm. The *hydrochloride* has m 280-285° from EtOH. [Meyer *Helv Chim Acta* 41 1890 1958, Beilstein 13 II 185, 13 III 856, 13 IV 889.]

**4-Amino-N,N-diethylaniline hydrochloride** [16713-15-8] M 200.7, m 233.5°, pK<sup>22</sup> 6.61. Crystallise the salt from EtOH. The *free base* [93-05-0] M 164.2 distils at 260-262°/~760mm. [Beilstein 14 IV 109.]

**4-Amino-3,5-diiodobenzoic acid** [2122-61-4] M 388.9, m ~350°, pK<sub>Est(1)</sub> 0.4, pK<sub>Est(2)</sub> ~1.6. Purify the iodo-acid by dissolving it in dilute NaOH and precipitating with dilute HCl. *Alternatively*, dissolve it in aqueous NH<sub>3</sub> and acidify it with AcOH. Dry it in air. The solubility of the *Na salt* in H<sub>2</sub>O is 2.56% at 25°. [Klemme &

Hunter *J Org Chem* **5** 510 1940, *Beilstein* **14** H 439, **14** III 1161, **14** IV 1284.]

**2-Aminodiphenylamine** [534-85-0] **M 184.2, m 79-80°, pK<sub>Est(1)</sub> ~3.8 (NH<sub>2</sub>), pK<sub>Est(2)</sub> <~0.** Crystallise the amine from H<sub>2</sub>O. [*Beilstein* **13** IV 43.]

**4-Aminodiphenylamine** [101-54-2] **M 184.2, b 155°/0.026mm, pK<sup>25</sup> 5.20.** It crystallises from EtOH with **m 66°**, and from ligroin with **m 75°**. It can be distilled at high vacuum. [*Beilstein* **13** IV 113.]

**2-Amino-1,2-diphenylethanol** [530-36-9] **M 213.3, m 165°, pK<sub>Est(1)</sub> ~7.5.** Recrystallise the ethanol from EtOH. The *1R,2S(-)*- [23190-16-1] and the *1S,2R(+)*- enantiomers also crystallise from EtOH and have **m 142-145°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 7° (c 0.6, EtOH).** [Masters et al. *J Org Chem* **56** 5666 1991, Masters & Hegedus *J Org Chem* **58** 4547 1993, *Beilstein* **13** IV 2150.]

**2-Aminodiphenylmethane (2-benzylaniline)** [28059-64-5] **M 183.3, m 52°, 51-54°, b 172°/12mm and 190°/22mm, pK<sub>Est(1)</sub> ~4.2.** Crystallise 2-benzylaniline from ether and distil it in a vacuum. [*Beilstein* **12** IV 3279.]

**2-Aminofluorene (2-fluorenamine)** [153-78-6] **M 181.2, m 127.8-128.8°, 132-133°, pK<sup>25</sup> 4.64.** Wash the amine well with H<sub>2</sub>O and recrystallise it from Et<sub>2</sub>O or 50% aqueous EtOH (25g with 400ml), and dry it in a vacuum. Store it in the dark. [Bavin *Org Synth Coll Vol V* 30 1973, *Beilstein* **12** H 1331, **12** IV 337.]

**9-Aminofluorene (9-fluorenamine)** [525-03-1] **M 181.2, m 64-65°, pK<sub>Est</sub> ~3.5.** Purify it by converting it to the hydrochloride with HCl, then basify it with NH<sub>3</sub> and recrystallise it from petroleum ether (**m 62-63°**) or from hexane. The *hydrochloride* [5978-75-6] **M 217.7**, has **m ~255°(dec, from EtOH).** [Ingold et al. *J Chem Soc* 1493 1933, Mathieu *Bull Soc Chim Fr* 1526 1971, *Beilstein* **12** H 1331, **12** I 553, **12** II 780, **12** III 3297, **12** IV 3390.]

**1-Amino-4-hydroxyanthraquinone** [116-85-8] **M 293.2, m 207-208°, pK<sub>Est(1)</sub> ~2.6 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~9.0 (OH).** Purify it by TLC on SiO<sub>2</sub> gel plates (0.75mm thick) using toluene/acetone (9:1) as eluent. The main band is scraped off and extracted with MeOH. The solvent is evaporated, and the dye is dried in a drying pistol [Land et al. *J Chem Soc, Faraday Trans 1* **72** 2091 1976]. It has also been recrystallised from aqueous EtOH. [*Beilstein* **14** H 268, **14** I 503, **14** II 168, **14** III 652, **14** IV 891.]

**3-Amino-4-hydroxytoluene (2-amino-*p*-cresol)** [95-84-1] **M 123.2, m 135°, 137-138°, pK<sub>Est(1)</sub> ~4.7(NH<sub>2</sub>), pK<sub>Est(2)</sub> ~9.6 (OH).** Recrystallise the cresol from H<sub>2</sub>O, Et<sub>2</sub>O or toluene. It sublimes *in vacuo* as plates or needles. The *hydrochloride* has **m 222-224° (dec, from aqueous EtOH).** [*Beilstein* **13** H 601, **13** I 227, **13** II 338, **13** III 1576.]

**4-Amino-3(5)-hydroxytoluene (2-amino-*m*-cresol)** [2835-98-5] **M 123.2, m 159°, 159-162°, pK<sub>Est(1)</sub> ~5.4 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~10.2 (OH).** Crystallise it from H<sub>2</sub>O, 50% EtOH, or toluene. [*Beilstein* **13** H 590, **13** III 1552.]

**2-Amino-5-hydroxytoluene** [2835-99-6] **M 123.2, m 162°(dec), 177-179°(dec), pK<sub>Est(1)</sub> ~5.4 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~10.4 (OH).** Crystallise it from 50% EtOH. [*Beilstein* **13** H 593, **13** IV 1698.]

**5-Aminoindane** [24425-40-9] **M 133.2, m 37-38°, b 131°/15mm, 146-147°/25mm, 247-249°/745mm, pK<sup>16</sup> 5.31.** Distil the indane, and then crystallise it from petroleum ether. [*Beilstein* **12** I 511, **12** III 2798.]

**1*RS*,2*SR*-(±)-*cis*-1-Amino-2-hydroxyindane (*cis*-1-amino-2-indanol)** [7480-35-5] **M 149.2, m 132-133°, 134-135°, pK<sub>Est</sub> ~8.5.** It is obtained by hydrolysis of (±)-*cis*-indano[1,2-*d*]-2-oxazolidone [in turn formed by solvolysis of ethyl *N*-(*trans*-2-iodo-1-indan)carbamate, **m 159.5-160°** in 87% yield by refluxing in dry diglyme for 14 hours]. The *cis*-2-oxazolidone (360mg) is dissolved in 0.85N methanolic KOH (20ml), diluted with H<sub>2</sub>O (5ml) and refluxed under N<sub>2</sub> for 18 hours, evaporated to dryness *in vacuo*, the residue is extracted with Et<sub>2</sub>O, filtered, dried (Mg<sub>2</sub>SO<sub>4</sub>), concentrated to a small volume and cooled to -12°, and the 2-*indanol* is collected as white plates (2 crops, **m 134-135° and 133-134°**) in 79% yield. [Hassner et al. *J Org Chem* **32** 540 1967.]

Similarly hydrolysis of *2-phenyl-3a,8a-dihydro-8H-indeno[1,2-d]oxazole* with diluted H<sub>2</sub>SO<sub>4</sub> (16 hours reflux), cool, basify with 10N NaOH, extract with CHCl<sub>3</sub>, dry (Mg<sub>2</sub>SO<sub>4</sub>), and evaporation gives an 88% yield of the (±)-*amino-alcohol* (m 132-133°). It can also be prepared from (±)-hydrindan-1,2-epoxide as described in the following entry. Its <sup>1</sup>HNMR (400MHz, CDCl<sub>3</sub>, TMS) has δ at 2.0-2.5 (br s, 3H), 2.95 (dd, *J* = 3 and 15Hz, 1H), 3.1 (dd, *J* = 7 and 15Hz, 1H), 4.28 (d, *J* = 7Hz, 1H), 4.4 (m, 1H), 7.2-7.35 (m, 4H). [Thompson et al. *J Med Chem* **35** 1685 1992.] Large scale preparations of the (±)-*cis*-1-amino-2-indanol from indene *via* the (±)-*trans*-acylamido-indanol by hydrolysis with strong acids (HCl, H<sub>2</sub>SO<sub>4</sub> or MeSO<sub>3</sub>H) which invert the OH stereochemistry have been patented [Gao et al. USP appl 5,599,985 1997; *Chem Abstr* **123** 55421b 1995, Gao et al. USP 5,616,808 1997, *Chem Abstr* **126** 250997h 1997]. [*Beilstein* **13** II 398a, 399d, **13** III 1841, **13** IV 2033.]

**1S,2R(-)-cis-1-Amino-2-hydroxyindane [(−)-cis-1-amino-2-indanol]** [126456-43-7] **M 149.2, m 117-117.5°, 118-121°, [α]<sub>D</sub><sup>20</sup> -61° (c 0.5, CHCl<sub>3</sub>), [α]<sub>D</sub><sup>25</sup> -62° (c 1.0, MeOH)**. This enantiomer has been prepared in several ways in high state of optical purity. Oxidation of indene [95-13-6] using Jacobsen's salen catalyst *R,R*-MnCl/NaOCl/CH<sub>2</sub>Cl<sub>2</sub> gave *1R,2S*-indane-1,2-epoxide in 84%ee, which on treatment with NH<sub>3</sub>/MeOH then PhCOCl/NaOH followed by 80% H<sub>2</sub>SO<sub>4</sub> provided *2-phenyl-3aS,8aR-dihydro-8H-indeno[1,2-d]oxazole* [see (±) above] by a C-1 chirality transfer process which gave the desired *1S,2R(-)*-aminoindanol on hydrolysis. Alternatively, oxidation of indene using Jacobsen's salen catalyst *S,S*-MnCl/PhCl/NaOCl and P<sub>3</sub>NO (see [34122-28-6]) gave the enantiomeric *1S,2R*-indane-1,2-epoxide which was reacted with Oleum/MeCN then hydrolysed with H<sub>2</sub>O (by C-2 chirality transfer *via* a Ritter reaction) to give the same desired *1S,2R(-)*-aminoindanol 78-80% yield with high optical purity. [Senanayake *Aldrichimica Acta* **31** 3 1998, Senanayake et al. *Tetrahedron Lett* **36** 3993 1995, Senanayake et al. *Tetrahedron Lett* **37** 3271 1996.] In a different approach *N*-(*L*-phenylalanyl)-(±)-amino-2-indanol was prepared [84% yield, from (±)-aminoindanol and BOC-*L*-phenylalanine *via* the 1-hydroxybenzotriazole/dimethyl 3-{3-(dimethylamino)-propyl}carbodiimide procedure], and the diastereoisomeric mixture was separated by low pressure chromatography on silica gel, eluting with 8% MeOH in CHCl<sub>3</sub> to give the faster moving *L*-Phe—*1S,2R*-1-aminoindan-2-ol diastereoisomer (40% yield). A mixture of this diastereomer (30g, 0.1mol), EtOH (1L) and 20% aqueous NaOH (265ml) was refluxed for 16 hours, concentrated to remove EtOH, diluted with H<sub>2</sub>O (100ml) and brine (100ml), extracted with CHCl<sub>3</sub> (3 x 600ml), the combined extracts were dried (Mg<sub>2</sub>SO<sub>4</sub>), filtered and evaporated to give *1S,2R*-1-aminoindan-2-ol (14g, 93%) as a white solid m 114-115°. Recrystallisation from Et<sub>2</sub>O/hexanes gave analytically pure white rods m 117-117.5°, with little loss of material, and with [α]<sub>D</sub><sup>25</sup> -62° (c 1.0, MeOH), with <sup>1</sup>HNMR (400MHz, CDCl<sub>3</sub>, TMS) which had δ at 2.0-2.5 (br s, 3H), 2.95 (dd, *J* = 3 and 15Hz, 1H), 3.1 (dd, *J* = 7 and 15Hz, 1H), 4.28 (d, *J* = 7Hz, 1H), 4.4 (m, 1H), 7.2-7.35 (m, 4H), same as the racemate. [Thompson et al. *J Med Chem* **35** 1685 1992.] Both enantiomers of this *cis*-aminoalcohol have also been obtained *via* enzymatic transformations [Didier et al. *Tetrahedron* **47** 4941 1991]. The oxazolines, oxazolidin-2-ones and the 2-*N*-tosylamides derived from the two enantiomers have been explored extensively as *chiral auxiliaries* [Senanayake *Aldrichimica Acta* **31** 3 1998]. For the preparation of CBS-oxazaboroline catalysis (Corey-Bakshi-Shibata) see Chapter 6, Catalysis-Part 1. **NOTE:** Both enantiomers are commercially available from Sepracore in gram to hundred kilogram quantities [Hong et al. *Tetrahedron Lett* **35** 6631 1994, see also Stinson *Chemical and Engineering News* May 16 p.6 1994]. [*Beilstein* **13** II 398a, 399d, **13** III 1841, **13** IV 2033.]

**1R,2S(+)-cis-1-Amino-2-hydroxyindane [(+)-cis-1-amino-2-indanol]** [136030-00-7] **M 149.2, m 118-121°, [α]<sub>D</sub><sup>20</sup> +63° (c 0.2, CHCl<sub>3</sub>)**. This compound was prepared using the same syntheses as for its enantiomer (preceding entry) and purified in the same way. It can be also obtained from the slower moving chromatographic band of the *L*-phenylalanyl diastereoisomer (see preceding entry).

**2-Amino-5-iodotoluene** [13194-68-8] **M 233.0, m 87°, pK<sub>Est</sub> ~3.6**. Crystallise it from 50% EtOH. [*Beilstein* **12** IV 1807.]

**1-Amino-4-methylaminoanthraquinone** [1220-94-6] **M 252.3, pK<sub>Est(1)</sub> ~1, pK<sub>Est(2)</sub> <~4**. Purify the quinone by TLC on silica gel plates using toluene/acetone (3:1) as eluent. The main band is scraped off and extracted with MeOH. The solvent is evaporated, and the residue is dried in a drying pistol [Land et al. *J Chem Soc, Faraday Trans 1* **72** 2091 1976]. [*Beilstein* **14** H 198, **14** I 462, **14** III 440, **14** IV 458.]

**4-Aminomethylbenzenesulfonamide hydrochloride (Mafenide HCl)** [138-37-4] **M 222.3, m 265-267°, pK<sub>1</sub><sup>20</sup>**

**8.18 (NH<sub>2</sub>), pK<sub>2</sub><sup>20</sup> 10.23 (SONH<sub>2</sub>).** Crystallise the salt from dilute HCl or 95% EtOH and dry it in a vacuum at 100°. It is an antibacterial. [Miller et al. *J Am Chem Soc* **62** 2099 1940, Bergeim & Barker *J Am Chem Soc* **66** 1459 1944.] The *sulfate salt* has **m** 254-255° (from aqueous EtOH). [*Beilstein* **14** III 2223, **14** IV 2799.]

**4-Amino-2-methyl-1-naphthol** See **Vitamin K<sub>5</sub>** in "Miscellaneous Compounds", Chapter 7.

**3-Amino-2-naphthoic acid** [5959-52-4] **M 187.2, m 214°(dec), pK<sub>Est(1)</sub> ~1.5 pK<sub>Est(2)</sub> ~4.0.** Crystallise the naphthoic acid from aqueous EtOH. [*Beilstein* **14** III 1341.]

**4-Amino-5-naphthol-2,7-disulfonic acid** [90-20-0] **M 320.3, pK<sub>1</sub><sup>20</sup> 3.54, pK<sub>2</sub><sup>20</sup> 8.55; pK<sub>1</sub><sup>25</sup> 3.63, pK<sub>2</sub><sup>25</sup> 8.83.** A slightly alkaline solution of Na<sub>2</sub>CO<sub>3</sub> (ca 22g) in water (litmus) is added to a solution of 100g of the dry acid in 750ml of hot distilled water, followed by 5g of activated charcoal and 5g of Celite. The suspension is stirred for 10 minutes and filtered by suction. The acid is then precipitated by adding ca 40ml of conc HCl (solution is blue to Congo Red), then filter it by suction through a shark skin filter circular sheet (or hardened filter paper) and wash it with 100ml of distilled water. The purification process is repeated. The acid is dried overnight in an oven at 60° and stored in a dark bottle. The *diethylamine salt* has **m** 306-307°(dec, from aqueous EtOH). [Post & Moore *Anal Chem* **31** 1872 1959]. [*Beilstein* **14** H 840, **14** I 758, **14** II 502, **14** III 2292, **14** IV 2824.]

**1-Amino-2-naphthol hydrochloride** [1198-27-2] **M 195.7, m 250°(dec), pK<sub>Est(1)</sub> ~3.7 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~9.9 (OH).** Crystallise the salt from the minimum volume of hot water containing a few drops of stannous chloride in an equal weight of hydrochloric acid (to reduce atmospheric oxidation). Filter the solution and add half its volume of conc HCl and set aside. The salt crystallises almost quantitatively. Dry it in a vacuum in the dark (**m** 260°). The salt is more stable than the *free base* which has **m** 150° (darkening at 130°) and its *O-methyl ether* has **m** 49° and **b** 159-159°/9mm. [Desai et al. *J Chem Soc* 324 1938, *Beilstein* **13** H 666, **13** I 268, **13** II 408, **13** III 1875, **13** IV 2080.]

**1-Amino-2-naphthol-4-sulfonic acid** [116-63-2] **M 239.3, m 295°(dec), pK<sub>Est(1)</sub> <0, pK<sub>Est(2)</sub> ~2.8 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~8.8.** Purify it by warming 15g of the acid, 150g of NaHSO<sub>3</sub> and 5g of Na<sub>2</sub>SO<sub>3</sub> (anhydrous) with 1L of water to ca 90°, shaking until most of the solid had dissolved, then filtering hot. The precipitate obtained by adding 10ml of conc HCl to the cooled filtrate is collected, washed with 95% EtOH until the washings are colourless, and is dried under vacuum over CaCl<sub>2</sub>. It is stored in a dark-coloured bottle, in the cold [Chanley et al. *J Am Chem Soc* **74** 4347 1952]. [*Beilstein* **14** IV 2825.]

**2-Amino-4-nitrobenzoic acid (4-nitroanthranilic acid)** [619-17-0] **M 182.1, m 269°(dec), pK<sub>1</sub><sup>25</sup> 0.65, pK<sub>2</sub><sup>25</sup> 3.70.** Crystallise the acid from water, EtOH (**m** 271°) or aqueous EtOH (**m** 269°). The *acetyl* derivative has **m** 217° (from EtOH), **m** 222° (from aqueous EtOH). [Chapman & Stephen *J Chem Soc* 1796 1925, *Beilstein* **14** II 234, **14** III 975, **14** IV 1087.]

**5-Amino-2-nitrobenzoic acid** [13280-60-9] **M 182.1, m 235°(dec), pK<sub>Est(1)</sub> ~1.1, pK<sub>Est(2)</sub> ~1.2.** Crystallise the acid from water. [*Beilstein* **14** III 1021.]

**1-Amino-4-nitronaphthalene** [776-34-1] **M 188.2, m 195°, 196-197°, pK<sup>20</sup> 0.54.** It crystallises from EtOH, ethyl acetate or aqueous NH<sub>3</sub> as light yellow crystals. The *acetyl* derivative also forms yellow crystals, **m** 192.5-193.5°, from Me<sub>2</sub>CO. [*Beilstein* **12** H 1259, **12** I 530, **12** II 704, **12** III 2971, **12** IV 3114.]

**2-Amino-4-nitrophenol** [99-57-0] **M 154.1, m 80-90° (hydrate), 142-143° (anhydrous), pK<sub>Est(1)</sub> ~3.9 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~9.2.** Crystallise the phenol from water. [*Beilstein* **13** IV 896.]

**2-Amino-5-nitrophenol** [121-88-0] **M 154.1, m 207-208°, pK<sub>Est(1)</sub> ~3.8, pK<sub>Est(2)</sub> ~9.3.** Crystallise the phenol from water. [*Beilstein* **13** IV 803.]

**RS-(±)-3-Amino-3-(4-nitrophenyl)propionic acid** [35005-61-9] **M 210.2, m 220-226°(dec), 226°(dec, browning at 215°), pK<sub>Est(1)</sub> ~3.0, pK<sub>Est(2)</sub> ~9.5.** The acid crystallises from 50% aqueous EtOH. The *hydrochloride* has **m** 218-220°(dec). The *N-benzoyl derivative* has **m** 204-205° (from MeOH) and the *N-*

*benzoyl-hydrazide* has **m** 226-227° (from MeOH or EtOH). [Posner *Justus Liebigs Ann Chem* **389** 44 1912, *Beilstein* **14** I 603, **14** IV 1548.]

**2-Aminophenol** [95-55-6] **M 109.1**, **m** 175-176°, **pK<sub>1</sub><sup>25</sup>** 4.65, **pK<sub>2</sub><sup>25</sup>** 9.75. Purify it by dissolving it in hot water, decolorising with activated charcoal, filtering and cooling to induce crystallisation. Maintain an atmosphere of N<sub>2</sub> over the hot phenol solution to prevent its oxidation [Charles & Freiser *J Am Chem Soc* **74** 1385 1952]. It can also be crystallised from EtOH using the same precautions. [*Beilstein* **13** IV 805.]

**3-Aminophenol** [591-27-5] **M 109.1**, **m** 122-123°, **pK<sub>1</sub><sup>25</sup>** 4.25, **pK<sub>2</sub><sup>25</sup>** 9.90. Crystallise it from hot water or toluene. [*Beilstein* **13** IV 952.]

**4-Aminophenol** [123-30-8] **M 109.1**, **m** 190° (under N<sub>2</sub>), **pK<sub>1</sub><sup>25</sup>** 5.38, **pK<sub>2</sub><sup>25</sup>** 10.4. Crystallise it from EtOH, then water, excluding oxygen. It sublimes at 110°/0.3mm. It has been purified by chromatography on alumina with a 1:4 (v/v) mixture of absolute EtOH/\*benzene as eluent. [*Beilstein* **13** IV 1014.]

**4-Aminophenol hydrochloride** [51-78-5] **M 145.6**, **m** 306°(dec). Purify the salt by treating an aqueous solution with saturated Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>, filtering under N<sub>2</sub>, then recrystallising it from 50% EtOH twice and once from absolute EtOH [Livingston & Ke *J Am Chem Soc* **72** 909 1950]. [*Beilstein* **13** III 993.]

**4-Aminophenylacetic acid** [1197-55-3] **M 151.2**, **m** 199-200°(dec), **pK<sub>1</sub><sup>20</sup>** 3.60, **pK<sub>2</sub><sup>20</sup>** 5.26. Crystallise the acid from hot water (60-70ml/g). [*Beilstein* **14** III 1182.]

**1RS,2RS-2-Amino-1-phenylbutan-1-ol** [ $\alpha$ -( $\alpha$ -aminopropyl)benzyl alcohol] [( $\pm$ )-*threo* 5897-76-7] **M 165.1**, **m** 79-80°, **pK<sub>Est</sub>** ~9.7. Crystallise the *free base* of the *threo* isomer from \*benzene/petroleum ether which has **m** 79-80°. The *threo-hydrochloride*, **m** 204-205°, crystallised from EtOH [Abrams & Kipping *J Chem Soc* 1480 1936, Rebstock et al. *J Am Chem Soc* **73** 3666 1951]. The *1RS,2RS-erythro* isomer has **m** 80.5-81° and its *hydrochloride* has **m** 242° [Harturg et al. *J Am Chem Soc* **52** 3317 1930. [*Beilstein* **13** II 390, **13** III 1791, **13** IV 1952.]

**N-Aminophthalimide** [1875-48-5] **M 162.2**, **m** 200-202°, **pK<sub>Est</sub>** ~0. The imide has been recrystallised from 96% EtOH (solubility is ~2% at boiling temperature) to form a yellow solution. It sublimes *in vacuo* at *ca* 150°. It solidifies after melting, and remelts at 338-341° (see phthalhydrazide “Heterocyclic Compounds” in this Chapter). [*Beilstein* **22/11** V 122.]

**4-Aminopropiophenone** [70-69-9] **M 163.1**, **m** 140°, **b** 180°/10mm, **pK<sub>Est</sub>** ~2.2. Crystallise it from water or aqueous EtOH. The *hydrochloride* has **m** 188-189° (aqueous EtOH/HCl drops), and the *semicarbazone* has **m** 139-140° (from EtOH/H<sub>2</sub>O). [Derrick & Bornemann *J Am Chem Soc* **35** 1283 1913, *Beilstein* **14** H 59, **14** I 375, **14** III 146, **14** IV 139.]

**4-(2-Aminopropyl)phenol (hydroxyamphetamine)** [103-86-6] **M 151.2**, **m** 125-126°, **pK<sub>Est(1)</sub>** ~9.4 (OH), **pK<sub>Est(2)</sub>** ~9.7(NH<sub>2</sub>), **pK<sub>2</sub><sup>25</sup>** 9.31. Crystallise the phenol from \*benzene. The *R(-)-enantiomer* crystallises from EtOH or \*C<sub>6</sub>H<sub>6</sub> with **m** 110.5-111.5° and  $[\alpha]_D^{17}$  -52.0° (c 1, EtOH). The ( $\pm$ )-*hydrochloride* has **m** 171-172° (EtOH/Et<sub>2</sub>O), and the *2,4-dinitrophenylhydrazone* has **m** 190-192° (EtOH). [*Beilstein* **13** I 251, **13** III 1709, **13** IV 1871.]

**1-Aminopyrene** [1606-67-3] **M 217.3**, **m** 117-118°, **pK<sub>1</sub><sup>25</sup>** 2.91 (50% aqueous EtOH), **pK<sub>2</sub><sup>25</sup>** 2.77 (50% aqueous EtOH). Crystallise it from hexane. [*Beilstein* **12** IV 3460.]

**4-Aminosalicyclic acid** [65-49-6] **M 153.1**, **m** 150-151°(dec), **pK<sub>1</sub><sup>25</sup>** 1.78 (CO<sub>2</sub>H), **pK<sub>2</sub><sup>25</sup>** 3.63 (NH<sub>2</sub>), **pK<sub>2</sub><sup>25</sup>** 13.74 (OH). Crystallise the acid from EtOH. [*Beilstein* **14** IV 1967.]

**5-Aminosalicyclic acid (5-amino-2-hydroxybenzoic acid)** [89-57-6] **M 153.1**, **m** 276-280°, 283°(dec), **pK<sub>1</sub><sup>25</sup>** 2.74 (CO<sub>2</sub>H), **pK<sub>2</sub><sup>25</sup>** 5.84 (NH<sub>2</sub>). It crystallises as needles from H<sub>2</sub>O containing a little NaHSO<sub>3</sub> to avoid aerial oxidation to the quinone-imine. The *Me ester* gives needles from \*C<sub>6</sub>H<sub>6</sub>, **m** 96°, and the *hydrazide* has **m** 180-

182° (from H<sub>2</sub>O). [Fallab et al. *Helv Chim Acta* **34** 26 1951, Shavel *J Amer Pharm Assoc* **42** 402 1953, *Beilstein* **14** IV 2058.]

**2-Amino-5-sulfanyllthiazole (thiazolsulfone)** [473-30-3] **M 238.3, m 219-221°(dec), pK<sub>Est</sub> ~4.5 (OH)**. Crystallise the thiazole from EtOH. It is antibacterial. [Bambas *J Am Chem Soc* **67** 671 1945.]

**4-Amino-2-sulfobenzoic acid** [527-76-4] **M 217.1**. Crystallise the sulfobenzoic acid from H<sub>2</sub>O (solubility is 0.3% at 25°). [KasHe *J Am Chem Soc* **44** 490 1910, *Beilstein* **19** I 356, **19** III/IV 1641 for 4-NO<sub>2</sub>.]

**9-Aminotriptycene** [793-41-9] **M 269.3, m 223.5-224.5°**. Recrystallise aminotriptycene from ligroin [Imashiro et al. *J Am Chem Soc* **109** 729 1987].

***p*-tert-Amylphenol (*p*-2,2-dimethylpropylphenol)** [80-46-6] **M 146.3, m 93.5-94.2°, 94-96°, b 138°/15mm, 262°/760mm, pK<sub>Est</sub> ~10.2**. Purify *via* its benzoate, as for phenol. After evaporating the solvent from its solution in ether, the material is recrystallised (from the melt) to a constant melting point. The *benzoyl* derivative has **m 60°** (from EtOH). [Berliner et al. *J Am Chem Soc* **76** 507 1954, Huston et al. *J Am Chem Soc* **67** 899 1945, *Beilstein* **6** H 548, **6** I 269, **6** II 506, **6** III 1965, **6** IV 3383.]

**Aniline** [62-53-3] **M 93.1, f -6.0°, b 68.3/10mm, 184.4°/760mm, d<sub>4</sub><sup>20</sup> 1.0220, n<sub>D</sub><sup>20</sup> 1.585, n<sub>D</sub><sup>25</sup> 1.5832, pK<sub>25</sub> 4.60**. Aniline is *hygroscopic*. It can be dried with KOH or CaH<sub>2</sub>, and distilled under reduced pressure. Treatment with stannous chloride removes sulfur-containing impurities, reducing the tendency to become coloured by aerial oxidation. It can be crystallised from Et<sub>2</sub>O at low temperatures. More extensive purifications involve preparation of derivatives, such as the double salt of aniline hydrochloride and cuprous chloride or zinc chloride, or *N*-acetylaniline (**m 114°**) which can be recrystallised from water.

Redistilled aniline is dropped slowly into a strong aqueous solution of recrystallised oxalic acid. Aniline oxalate (**m 174-175°**) is filtered off, washed several times with water and recrystallised three times from 95% EtOH. Treatment with saturated Na<sub>2</sub>CO<sub>3</sub> solution regenerated aniline which was distilled from the solution, dried and redistilled under reduced pressure [Knowles *Ind Eng Chem* **12** 881 1920].

After refluxing with 10% acetone for 10 hours, aniline is acidified with HCl (Congo Red as indicator) and extracted with Et<sub>2</sub>O until colourless. The hydrochloride is purified by repeated crystallisation before aniline is liberated by addition of alkali, then dried with solid KOH, and distilled. The product is sulfur-free and remains colourless in air [Hantzsch & Freese *Chem Ber* **27** 2529, 2966 1894].

Non-basic materials, including nitro compounds, are removed from aniline in 40% H<sub>2</sub>SO<sub>4</sub> by passing steam through the solution for 1 hour. Pellets of KOH are then added to liberate the aniline which is steam distilled, dried with KOH, distilled twice from zinc dust at 20mm, dried with freshly prepared BaO, and finally distilled from BaO in an all-glass apparatus [Few & Smith *J Chem Soc* 753 1949]. Aniline is absorbed through skin and is **TOXIC**. [*Beilstein* **12** IV 223.]

**Aniline hydrobromide** [542-11-0] **M 174.0, m 286°**. Crystallise the hydrobromide from water or EtOH and dry it at 5mm over P<sub>2</sub>O<sub>5</sub>. Also recrystallise it four times from MeOH containing a few drops of conc HBr by addition of petroleum ether (b 60-70°), then dry it to constant weight over paraffin chips *in vacuo* [Gutbezahl & Grunwald *J Am Chem Soc* **75** 559 1953]. It precipitates from EtOH solution on addition of Et<sub>2</sub>O, and the filtered solid is recrystallised from EtOH and dried *in vacuo*. [Buchanan et al. *J Am Chem Soc* **108** 1537 1986, *Beilstein* **12** III 232.]

**Aniline hydrochloride** [142-04-1] **M 129.6, m 200.5-201°**. Purification is as for aniline HBr above. [*Beilstein* **12** IV 232.]

**Aniline hydroiodide** [45497-73-2] **M 220.0, it decomposes on heating**. Purification is as for aniline HBr; store it in the dark. [*Beilstein* **12** III 232.]

***m*-Anisaldehyde (3-methoxybenzaldehyde)** [591-31-1] **M 136.2, b 143°/50mm, d<sub>4</sub><sup>20</sup> 1.119**. Wash it with saturated aqueous NaHCO<sub>3</sub>, then H<sub>2</sub>O, dry it with anhydrous MgSO<sub>4</sub> and distil it under reduced pressure under N<sub>2</sub>. Store it under N<sub>2</sub> in sealed glass ampoules. [*Beilstein* **8** IV 241.]

***p*-Anisaldehyde (4-methoxybenzaldehyde)** [123-11-5] **M 136.2, m -1°, b 249°/atm, 89-90°/2mm, d<sub>4</sub><sup>20</sup> 1.119, n<sub>D</sub><sup>20</sup> 1.576.** Wash the aldehyde with saturated aqueous NaHCO<sub>3</sub>, then H<sub>2</sub>O, steam distil, extract the distillate with Et<sub>2</sub>O, dry (MgSO<sub>4</sub>) the extract, filter and distil this under a vacuum and N<sub>2</sub>. Store it in glass ampules under N<sub>2</sub> in the dark. [Beilstein 8 IV 252.]

***o*-Anisidine (2-methoxyaniline)** [90-04-0] **M 123.2, m ~5°, b 109°/17mm, 119°/21mm, 225°/atm, d<sub>4</sub><sup>20</sup> 1.096, n<sub>D</sub><sup>20</sup> 1.575, pK<sup>25</sup> 4.52.** It is separated from the *m*- and *p*- isomers by steam distillation. It is also separated from its usual synthetic precursor *o*-nitroanisole by dissolving it in dilute HCl (pH <2.0) extracting the nitro impurity with Et<sub>2</sub>O, adjusting the pH to ~8.0 with NaOH, extracting the amine into Et<sub>2</sub>O or steam distilling. Extract the distillate with Et<sub>2</sub>O, dry the extract (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and fractionate the residual oil. Protect the almost colourless oil from light which turns it yellow in color. [Biggs & Robinson *J Chem Soc* 388/1961, Nodzu et al. *Yakugaku Zasshi (J Pharm Soc Japan)* 71 713, 715 1951, Beilstein 13 IV 806.]

***m*-Anisidine (3-methoxyaniline)** [536-90-3] **M 123.2, m ~5°, b 79°/1mm, 128°/17mm, 251°/atm, d<sub>4</sub><sup>20</sup> 1.101, n<sub>D</sub><sup>20</sup> 1.583, pK<sup>25</sup> 4.20.** *o*-Isomer impurity can be removed by steam distillation. Another possible impurity is the precursor 3-nitroanisole which can be removed as for the preceding *o*-isomer and fractionating using an efficient column. It is a yellow liquid. [Gilman & Kyle *J Am Chem Soc* 74 3027 1952, Bryson *J Am Chem Soc* 82 4858 1960, Kadaba & Massie *J Org Chem* 22 333 1957, Beilstein 13 IV 953.]

***p*-Anisidine (4-methoxyaniline)** [104-94-9] **M 123.2, m 57°, pK<sup>25</sup> 5.31.** Crystallise *p*-anisidine from H<sub>2</sub>O or aqueous EtOH. Dry it in a vacuum oven at 35° for 6 hours and store it in a dry box. [More et al. *J Am Chem Soc* 108 2257 1986.] Purify it also by vacuum sublimation [Guarr et al. *J Am Chem Soc* 107 5104 1985]. [Beilstein 13 IV 1015.]

**Anisole** [100-66-3] **M 108.1, f -37.5°, b 43°/11mm, 153.8°/760mm, d<sub>15</sub><sup>15</sup> 0.9988, n<sub>D</sub><sup>20</sup> 1.5143, pK<sup>0</sup> -6.61 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Shake anisole with half its volume of 2M NaOH, and the emulsion is allowed to separate. Repeat three times, then wash twice with water, dry over CaCl<sub>2</sub>, filter, dry over sodium wire and finally distil it from fresh sodium under N<sub>2</sub> using a Dean-Stark trap (samples in the trap being rejected until free from turbidity) [Caldin et al. *J Chem Soc, Faraday Trans 1* 72 1856 1976]. Alternatively, dry it with CaSO<sub>4</sub> or CaCl<sub>2</sub>, or by refluxing with sodium or BaO with crystalline FeSO<sub>4</sub> or by passage through an alumina column. Traces of phenols are removed by prior shaking with 2M NaOH, followed by washing with water. It has been purified by zone refining. [Beilstein 6 IV 548.]

**2-*p*-Anisyl-1,3-indanone (anisindione)** [117-37-3] **M 252.3, m 156-157°, pK<sup>20</sup> 4.09.** Crystallise anisidinone from acetic acid or EtOH. [Horeau & Jacqius *Bull Soc Chim Fr* 53 1948, Koelsch *J Am Chem. Soc* 58 1331 1936, Beilstein 8 III 2931.]

**Anthracene** [120-12-7] **M 178.2, m 215-216°, 218°, b 342°/760mm, pK<sup>25</sup> -7.4 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Likely impurities are anthraquinone, anthrone, carbazole, fluorene, 9,10-dihydroanthracene, tetracene and bianthryl. Carbazole is removed by continuous-adsorption chromatography [see Sangster & Irvine *J Phys Chem* 24 670 1956] using a neutral alumina column and eluting with *n*-hexane. [Sherwood in *Purification of Inorganic and Organic Materials*, Zief (ed), Marcel Dekker, New York, 1969.] The solvent is evaporated, and anthracene is sublimed under vacuum, then purified by zone refining, under N<sub>2</sub> in darkness or non-actinic light.

It has also been purified by co-distillation with ethylene glycol (boils at 197.5°), from which it can be recovered by addition of water, followed by crystallisation from 95% EtOH, \*benzene, toluene, a mixture of \*benzene/xylene (4:1), or Et<sub>2</sub>O. It has also been chromatographed on alumina with petroleum ether in a dark room (to avoid photo-oxidation of adsorbed anthracene to anthraquinone). Other purification methods include sublimation in a N<sub>2</sub> atmosphere (in some cases after refluxing with sodium), and recrystallisation from toluene [Gorman et al. *J Am Chem Soc* 107 4404 1985].

Anthracene has been crystallised from EtOH, chromatographed through alumina in hot \*benzene (*fume hood*) and then sublimed in a vacuum in a pyrex tube that has been cleaned and baked at 100°. (For further details see Craig & Rajikan *J Chem Soc, Faraday Trans 1* 74 292 1978, and Williams & Zboinski *J Chem Soc, Faraday Trans 1* 74 611 1978.) It has been chromatographed on alumina, recrystallised from *n*-hexane and sublimed under reduced pressure. [Saltiel *J Am Chem Soc* 108 2674 1986, Masnori et al. *J Am Chem Soc* 108 1126 1986.]

*Alternatively*, recrystallise it from cyclohexane, chromatograph it on alumina with *n*-hexane as eluent, and recrystallise two more times [Saltiel et al. *J Am Chem Soc* **109** 1209 1987]. Anthracene is fluorescent and forms a *picrate complex*, **m** 139°, on mixing the components in CHCl<sub>3</sub> or \*C<sub>6</sub>H<sub>6</sub>, but decomposes on attempted crystallisation. [*Beilstein* **5** IV 228.]

**Anthracene-9-carboxylic acid (anthroic acid)** [723-62-6] **M** 222.2, **m** 214°(dec), **pK**<sup>20</sup> 3.65. Crystallise the acid from EtOH. It is fluorescent in EtOH. [*Beilstein* **9** IV 2671.]

**9-Anthraldehyde** [642-31-9] **M** 206.2, **m** 104-105°. Crystallise the aldehyde from acetic acid or EtOH. [Masnori et al. *J Am Chem Soc* **108** 1126 1986, *Beilstein* **7** IV 1738.]

**Anthranol (enol tautomer of anthrone, see below)** [529-86-2] **M** 196.2, **m** 120°(rapid heating), 160-170°(dec. but sainters >100°), **pK**<sup>25</sup> -5.5 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise anthrol from glacial acetic acid or aqueous EtOH. It is the less stable of two tautomers and is obtained by heating anthrone in aqueous NaOH (but not in cold NaOH) whereby it dissolves to form the sodium salt of anthrol. On cautious acidification anthrol precipitates as a yellowish-brown solid. It changes into anthrone on keeping, as it does on melting (120° in a preheated bath). The melting point depends on the initial bath temperature, and recrystallisation may cause it to tautomerise to *anthrone* (see below). The *acetate* (**m** 134°) crystallises from petroleum ether (**m** 134°) or EtOH (**m** 135.5-137°). [Meyer *Justus Liebigs Ann Chem* **379** 56 1911, Barnett et al. *J Chem Soc* 885 1928, for tautomerism see Almdal et al. *Acta Chem Scand Series B* **40** 230 1986.]

**Anthranthrone** [641-13-4] **M** 306.3, **m** 300°, 415°, **pK**<sup>25</sup> -7.9 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise it from chlorobenzene, nitrobenzene or CHCl<sub>3</sub> (**m** 340°). [*Beilstein* **7** I 451, **7** II 783, **7** III 4406, **7** IV 2694.]

**Anthraquinone** [84-65-1] **M** 208.2, **m** 286°, **pK**<sup>25</sup> -8.27 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise anthraquinone from CHCl<sub>3</sub> (38ml/g), \*benzene, or boiling acetic acid, wash it with a little EtOH and dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **7** IV 2556.]

**Anthrarufin (1,5-dihydroxy-9,10-anthraquinone)** [117-12-4] **M** 240.1, **m** 280°(dec), **pK**<sub>1</sub><sup>25</sup> 9.90, **pK**<sub>2</sub><sup>25</sup> 11.05. Purify anthrarufin by column chromatography on silica gel with CHCl<sub>3</sub>/Et<sub>2</sub>O as eluent, followed by recrystallisation from acetone. *Alternatively*, recrystallise it from glacial acetic acid [Flom & Barbara *J Phys Chem* **89** 4489 1985]. [*Beilstein* **7** III 2359, **8** IV 3268.]

**1,8,9-Anthratriol** [480-22-8] **M** 226.2, **m** 176-181°, **pK**<sub>Est</sub> ~9.5. Crystallise it from petroleum ether and dry it *in vacuo*. [*Beilstein* **6** IV 7602.]

**Anthrimide (1,1'-imino-bis-anthraquinone)** [82-22-4] **M** 429.4, **m** >250°(dec). Crystallise anthrimide from chlorobenzene (red needles) or nitrobenzene (red rhombs). It is poorly soluble in organic solvents. [Eckert & Steiner *Monatsh Chem* **35** 1129 1914.]

**Anthrone** [90-44-8] **M** 194.2, **m** 155°, **pK**<sup>25</sup> -5.5 (aqueous H<sub>2</sub>SO<sub>4</sub>). This stable keto tautomer of 9-anthranol (above) provides yellow crystals from a 3:1 mixture of \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 60-80°) (10-12ml/g), or successively from \*C<sub>6</sub>H<sub>6</sub> then EtOH. Dry it *in vacuo*. [Meyer *Org Synth Coll Vol I* 60 1941, *Beilstein* **6** IV 4930.]

(±)-**Atrolactic acid (0.5H<sub>2</sub>O) (2-hydroxy-2-phenylpropionic acid)** [515-30-0] **M** 166.2, **m** 94.5-95°(anhydrous), 88-91° (0.5H<sub>2</sub>O), **pK**<sup>18</sup> 3.53. Crystallise the acid from water (4g/20ml boiling H<sub>2</sub>O with charcoal, filter and cool overnight at 0-5°) and dry it at 55°/0.5mm [Eliel & Freeman *Org Synth Coll Vol IV* 58 1963]. The (±)-*ethyl ester* [76496-51-0] has **b** 250°/~760mm, 129-130°/13mm, and the (±)-*amide* [5908-94-1] has **m** 101-102°. [*Beilstein* **10** H 259, **10** I 113, **10** II 157, **10** III 560, **10** IV 467.]

**R(-) and S(+)** **Atrolactic acid** [*R* 3966-30-1 and *S* 13113-71-8] **M** 166.2, **m** 115-116°, 116-117°, **R** [ $\alpha$ ]<sub>D</sub><sup>14.6</sup> -37.7° (c 3.3 EtOH) and [ $\alpha$ ]<sub>D</sub><sup>14.6</sup> -51.1° (c 2.2, H<sub>2</sub>O), **S** [ $\alpha$ ]<sub>D</sub><sup>14.6</sup> +37.7° (c 3.5 EtOH). Recrystallise them from \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and dry them in a vacuum. [McKenzie & Clough *J Chem Soc* 1019 1910,

Meyers & Slade *Synth Commun* **6** 601/1972, *Beilstein* **10** II 157, **10** III 560, **10** IV 657.]

**Auramine O** (4,4'-bis-dimethylaminobenzophenone imine hydrochloride) [2465-27-2] **M 321.9, m >250°(dec), pK<sup>25</sup> 10.71 (free base), 9.78 (carbinolamine)**. It crystallises from EtOH as the hydrochloride and is very slightly soluble in CHCl<sub>3</sub>, its UV has  $\lambda_{\max}$  at 434nm ( $\epsilon$  370). The *free base* (carbinolamine) has **m 136°** (from \*C<sub>6</sub>H<sub>6</sub>). [Goldacre & Phillips *J Chem Soc* 1724 1949, Conrad et al. *Biochemistry* **9** 1540 1970, *Beilstein* **14** IV 256.]

**Aurin tricarboxylic acid** [4431-00-9] **M 422.4, m 300°**. The acid is dissolved in aqueous NaOH, NaHSO<sub>3</sub> solution is added until the colour is discharged and then the tricarboxylic acid is precipitated with HCl. [Heisig & Lauer *Org Synth Coll Vol I* 54 1941, *Beilstein* **10** IV 4161]. Do not extract the acid with hot water because it softens, forming a viscous mass. Make a solution in aqueous NH<sub>3</sub>. **Aluminon** is the NH<sub>4</sub> salt.

**Azobenzene** [103-33-3] **M 182.2, m 68°, b 293°/atm, pK<sup>25</sup> 2.48**. Ordinary azobenzene is nearly all in the *trans*-form. It is partly converted into the *cis*-form on exposure to light [for isolation see Hartley *J Chem Soc* 633 1938, and for spectra of *cis*- and *trans*-azobenzenes, see Winkel & Siebert *Chem Ber* **74B** 670/1941]. *trans*-Azobenzene is obtained by chromatography on alumina using 1:4 \*C<sub>6</sub>H<sub>6</sub>/heptane or petroleum ether, and it crystallises from EtOH (after refluxing for several hours) or hexane. All operations should be carried out in diffuse red light or in the dark. [*Beilstein* **16** IV 8.]

**4-Azidoaniline hydrochloride** [91159-79-4] **M 170.6, m 165°(dec)**. Purify it in EtOAc with dry HCl gas followed by cold dry Et<sub>2</sub>O. The *free base* has **m 66°(dec)** (MeOH) and the *picrate* has **m 64-65°(dec)** (MeOH). The IR has  $\nu_{\max}$  at (Nujol) at 2110cm<sup>-1</sup> (N<sub>3</sub>). [Escher et al. *Helv Chim Acta* **62** 1217 1979, Maffei & Rivolta *Gazetta Chim Ital* **84** 750 1954, *Beilstein* **12** H 772, **12** IV 1741.]

**Azolitmin B** [1395-18-2] **M ~3300, m >250°(dec)**. It crystallises from water as dark violet scales, or as a red powder on precipitation from H<sub>2</sub>O by addition of EtOH. It is an indicator which is red at pH 4.5 and blue at pH 8.3. [cf Kolthopff & Rosenblum *Acid-Base Indicators*, Macmillan, NY pp160-162, 365-366 1937.]

*p,p'*-**Azoxyanisole** (4,4'-dimethoxyazoxybenzene) [1562-94-3] **M 258.3, transition temperatures: 118.1-118.8°, 135.6-136.0°, pK<sup>25</sup> -5.23 (20% aqueous EtOH + 80% aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise the dye from absolute or 95% EtOH, or acetone, and dry it by heating under vacuum or sublime it in a vacuum onto a cold finger. [*Beilstein* **16** II 326.]

**Azoxybenzene (Fenazox)** [495-48-7] **M 198.2, m 36°, pK<sup>25</sup> -6.16 (20% aqueous EtOH + 80% aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise azobenzene from EtOH or MeOH, and dry it for 4 hours at 25°/10<sup>-3</sup>mm. Sublime it before use. [Bigelow & Palm *Org Synth Coll Vol II* 57 1943, *Beilstein* **16** II 326.]

*p,p'*-**Azoxyphenetole** [4792-83-0] **M 286.3, m 137-138° (turbid liquid clarifies at 167°)**. Crystallise the dye from toluene or EtOH. [*Beilstein* **16** II 326.]

**Azulene** [275-51-4] **M 128.2, m 98.5-99°, pK<sup>25</sup> -1.65 (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise azulene from EtOH. It has UV with  $\lambda_{\max}$  at 270nm (log $\epsilon$  4.72) in hexane. [Platner & Magyar *Helv Chim Acta* **25** 581 1942, *Beilstein* **5** IV 1636.]

**Benzalacetone** (*trans*-4-phenyl-3-buten-2-one) [122-57-6] **M 146.2, m 42°**. Crystallise it from petroleum ether (b 40-60°), or distil it (b 137-142°/16mm). [*Beilstein* **7** IV 1003.]

**Benzalacetophenone (Chalcone)** [94-41-7] **M 208.3, m 56-58°, b 208°/25mm, pK<sup>25</sup> -5.73 (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise it from EtOH by warming to 50° (about 5ml/g), iso-octane, or toluene/petroleum ether, or recrystallise it from MeOH, and then twice from hexane. SKIN IRRITANT. [*Beilstein* **7** IV 1658.]

**Benzaldehyde** [100-52-7] **M 106.1, f -26°, b 62° 58°/10mm, 179.0°/760mm, d<sub>4</sub><sup>20</sup> 1.044, n<sub>D</sub><sup>20</sup> 1.5455, pK<sup>25</sup> -7.1**

**(aqueous H<sub>2</sub>SO<sub>4</sub>).** To diminish its rate of oxidation, benzaldehyde usually contains additives such as hydroquinone or catechol. It can be purified *via* its bisulfite addition compound but usually distillation (under nitrogen at reduced pressure) is sufficient. Prior to distillation it is washed with NaOH or 10% Na<sub>2</sub>CO<sub>3</sub> (until no more CO<sub>2</sub> is evolved), then with saturated Na<sub>2</sub>SO<sub>3</sub> and H<sub>2</sub>O, followed by drying with CaSO<sub>4</sub>, MgSO<sub>4</sub> or CaCl<sub>2</sub>. [Beilstein 7 IV 505.]

**anti-Benzaldoxime** [932-90-1] M 121.1, m 33-34°. Crystallise the oxime from diethyl ether by adding petroleum ether (b 60-80°). The *syn*-isomer [622-32-2] has b 121-124°/12mm, m 34-36°. [Beilstein 7 H 218, 7 IV 527.]

**Benzamide** [55-21-0] M 121.1, m 129.5°, pK<sup>25</sup> -2.16 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise it from hot water (about 5ml/g), EtOH or 1,2-dichloroethane, and dry it in air. It has also been crystallised from dilute aqueous NH<sub>3</sub>, H<sub>2</sub>O, Me<sub>2</sub>CO, then \*C<sub>6</sub>H<sub>6</sub> using a Soxhlet extractor. Dry it in an oven at 110° for 8 hours and store in a desiccator over 99% H<sub>2</sub>SO<sub>4</sub>. [Bates & Hobbs *J Am Chem Soc* 73 2151 1951, Beilstein 9 IV 725.]

**Benzamidine** [618-39-3] M 120.2, m 64-66°, pK<sup>20</sup> 11.6. It is liberated from its hydrochloride chloride (below) by treatment with 5M NaOH, extracted into diethyl ether, dried (Na<sub>2</sub>SO<sub>4</sub>) and sublimed *in vacuo*. [Beilstein 9 H 280, 9 I 123, 9 II 199, 9 IV 1264, 9 IV 898.]

**Benzamidine hydrochloride hydrate** [206752-36-5] M 156.6 (anhydrous), m 86-88°, pK<sup>20</sup> 11.6 (see free base above). Recrystallise it from dilute HCl (crystals contain xH<sub>2</sub>O) or EtOH/few drops HCl; dry it in a vacuum. It is a proteolytic inhibitor [Jeffcoate & White *J Clin Endocrinol Metab* 38 155 1974, Beilstein 9 IV 898.]

**Benzanilide** [93-98-1] M 197.2, m 164°, pK<sup>55</sup> 1.26. Crystallise benzanilide from petroleum ether (b 70-90°) using a Soxhlet extractor, and dry it overnight at 120°. Also crystallise it from EtOH. [Beilstein 12 IV 417.]

**Benz[a]anthracene (1,2-benzanthracene, tetraphene)** [56-55-3] M 228.3, m 159-160°. Crystallise 1,2-benzanthracene from MeOH, EtOH or \*benzene (charcoal), then chromatograph it on alumina from sodium-dried \*benzene (twice), using vacuum distillation to remove \*benzene. Final purification is by vacuum sublimation. [Beilstein 5 IV 2549.]

**Benz[a]anthracene-7,12-dione** [2498-66-0] M 258.3, m 169.5-170.5°, 169-171°. Crystallise the dione from MeOH (charcoal), toluene, toluene/hexane, Me<sub>2</sub>CO, or AcOH. Alternatively purify it by sublimation *in vacuo*, or by zone refining. [Beilstein 7 H 826, 7 I 440, 7 II 760, 7 III 4278, 7 IV 2644.]

**Benzanthrone** [82-05-3] M 230.3, m 170°, pK<sup>25</sup> -3.2 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise benzanthrone from EtOH or xylene. [Beilstein 7 IV 1819.]

**\*Benzene** [71-43-2] M 78.1, f 5.5°, b 80.1°, d<sub>4</sub><sup>20</sup> 0.874, n<sub>D</sub><sup>20</sup> 1.50110, n<sub>D</sub><sup>25</sup> 1.49790. For most purposes, \*benzene can be purified sufficiently by shaking with conc H<sub>2</sub>SO<sub>4</sub> until free from thiophene, then with H<sub>2</sub>O, dilute NaOH and water, followed by drying (with P<sub>2</sub>O<sub>5</sub>, sodium, LiAlH<sub>4</sub>, CaH<sub>2</sub>, 4X Linde molecular sieve, or CaSO<sub>4</sub>, or by passage through a column of silica gel, and for a preliminary drying, CaCl<sub>2</sub> is suitable), and distillation. A further purification step to remove thiophene, acetic acid and propionic acid, is crystallisation by partial freezing. The usual contaminants in dry thiophene-free \*benzene are non-benzenoid hydrocarbons such as cyclohexane, methylcyclohexane, and heptanes, together with naphthenic hydrocarbons and traces of toluene. Carbonyl-containing impurities can be removed by percolation through a *Celite column impregnated with 2,4-dinitrophenylhydrazine, phosphoric acid and H<sub>2</sub>O*. (Prepared by dissolving 0.5g DNPH in 6ml of 85% H<sub>3</sub>PO<sub>4</sub> by grinding together, then adding and mixing 4ml of distilled H<sub>2</sub>O and 10g Celite.) [Schwartz & Parker *Anal Chem* 33 1396 1961.] \*Benzene has been freed from thiophene by refluxing with 10% (w/v) of Raney nickel for 15 minutes, after which the nickel is removed by filtration or centrifugation.

Dry \*benzene is obtained by doubly distilling high purity \*benzene from a solution containing the blue ketyl formed by the reaction of sodium-potassium alloy with a small amount of benzophenone.

Thiophene has been removed from \*benzene (absence of a bluish-green coloration when 3ml of \*benzene is

shaken with a solution of 10mg of isatin in 10ml of conc H<sub>2</sub>SO<sub>4</sub>) by refluxing the \*benzene (1.25L) for several hours with 40g HgO (freshly precipitated) dissolved in 40ml glacial acetic acid and 300ml of water. The precipitate is filtered off, the aqueous phase is removed and the \*benzene is washed twice with H<sub>2</sub>O, dried and distilled. *Alternatively*, \*benzene dried with CaCl<sub>2</sub> has been shaken vigorously for 0.5 hour with anhydrous AlCl<sub>3</sub> (12g/L) at 25-35°, then decanted, washed with 10% NaOH, and water, dried and distilled. The process is repeated, giving thiophene-free \*benzene. [Holmes & Beeman *Ind Eng Chem* **26** 172 1934.]

After shaking successively for about an hour with conc H<sub>2</sub>SO<sub>4</sub>, distilled water (twice), 6M NaOH, and distilled water (twice), \*benzene is distilled through a 3-ft glass column to remove most of the water. Absolute EtOH is added and the \*benzene-alcohol azeotrope is distilled. (This low-boiling distillation leaves any non-azeotrope-forming impurities behind.) The middle fraction is shaken with distilled water to remove EtOH, and again redistilled. Final slow and very careful fractional distillation from sodium, then LiAlH<sub>4</sub> under N<sub>2</sub>, removed traces of water and peroxides. [Peebles et al. *J Am Chem Soc* **82** 2780 1960.] \*Benzene liquid and vapour are very **TOXIC, CARCINOGENIC and HIGHLY FLAMMABLE**, and all operations should be carried out in an efficient fume cupboard and in the absence of naked flames in the vicinity. [Beilstein **5** H 175, **5** I 95, **5** II 119, **5** III 469.]

**Rapid purification:** To dry benzene, alumina, CaH<sub>2</sub> or 4A molecular sieves (3% w/v) may be used (dry for 6 hours). Then benzene is distilled, discarding the first 5% of distillate, and stored over molecular sieves (3A, 4A) or Na wire.

[<sup>2</sup>H<sub>6</sub>]\*Benzene (\*benzene-d<sub>6</sub>) [1076-43-3] **M 84.2, b 80°/773.6mm, 70°/562mm, 60°/399mm, 40°/186.3mm, 20°/77.1mm, 10°/49.9mm, 0°/27.5mm, d<sub>4</sub><sup>20</sup> 0.9488, d<sub>4</sub><sup>40</sup> 0.9257, n<sub>D</sub><sup>20</sup> 1.4991, n<sub>D</sub><sup>40</sup> 1.4865.** Hexadeuteriobenzene of 99.5% purity is refluxed over and distilled from CaH<sub>2</sub> onto Linde type 5A sieves under N<sub>2</sub>. [Beilstein **5** III 518, **5** IV 630.]

**Benzeneazodiphenylamine (4-phenylazodiphenylamine)** [28110-26-1; 101-75-7] **M 273.3, m 82°, 86°, 87-91°, pK<sup>22</sup> 0.48.** Purify the dye by chromatography on neutral alumina using dry \*C<sub>6</sub>H<sub>6</sub> with 1% of dry MeOH. The major component, which gave a stationary band, is cut out and eluted with EtOH or MeOH. [Högfeldt & Bigeleisen *J Am Chem Soc* **82** 15 1960.] It crystallises from petroleum ether, EtOH or aqueous EtOH, and has λ<sub>max</sub> at 420nm (ε 28,000) (aqueous EtOH) and 540nm (aqueous EtOH/H<sub>2</sub>SO<sub>4</sub>) [Badger et al. *J Chem Soc* 1888 1954, Beilstein **16** H 314, **16** III 343, **16** IV 457.]

**1-Benzeneazo-2-naphthol (Sudan I)** [842-07-9] **M 248.3, m 106°, 120.5°, 132° (polymorphism) 134°, 135°, pK<sub>Est</sub> ~9.5 (OH).** Crystallise the dye from EtOH. It forms Cu and Ni salts. [Beilstein **16** H 162, **16** I 254, **16** II 70, **16** III 129, **16** IV 228.]

**1-Benzeneazo-2-naphthylamine (1-phenylazo-2-naphthylamine, Yellow AB)** [85-84-7] **M 247.3, m 102-104°, 103-104°, pK<sub>Est</sub> ~4.1.** Crystallise the dye from glacial acetic acid, acetic acid/water or ethanol. It forms Cu, Co and Ni salts. [Beilstein **16** H 369, **16** I 328, **16** II 193, **16** III 417, **16** IV 551.]

**1,2-Benzenedimethanol (1,2-bishydroxymethylbenzene, phthalyl alcohol)** [612-14-6] **M 138.2, m 61-64°, 63-64°, 64-65°, 65-66.5°, b 145°/3mm.** Recrystallise it from \*C<sub>6</sub>H<sub>6</sub>, H<sub>2</sub>O, petroleum ether or pentane. It has been extracted in a Soxhlet with Et<sub>2</sub>O, evaporated and recrystallised from hot petroleum ether. It is also purified by dissolving in Et<sub>2</sub>O, allowing to evaporate till crystals are formed, filtering off and washing the colourless crystals with warm petroleum ether or pentane. The *diacetate* has **m 35°, 35-36°.** [Nystrom & Brown *J Am Chem Soc* **69** 1197 1947, IR and UV: Entel et al. *J Am Chem Soc* **74** 441 1952, Beilstein **6** IV 5953.]

**m-Benzenedisulfonic acid** [98-48-6] **M 238.2, pK<sub>Est</sub> <0.** Free it from H<sub>2</sub>SO<sub>4</sub> by conversion to the calcium or barium salts (using Ca(OH)<sub>2</sub> or Ba(OH)<sub>2</sub>, and filtering). The calcium salt is then converted to the potassium salt, using K<sub>2</sub>CO<sub>3</sub>. Both the potassium and the barium salts are recrystallised from H<sub>2</sub>O, and the acid is regenerated by passing through the H<sup>+</sup> form of a strong cation exchange resin. The acid is recrystallised twice from conductivity water and dried over CaCl<sub>2</sub> at 25°. [Atkinson et al. *J Am Chem Soc* **83** 1570 1961.] It has also been crystallised from Et<sub>2</sub>O and dried in a vacuum oven. The *S-benzylisothiuronium salt* has **m 214.3°** (from EtOH/H<sub>2</sub>O). [Beilstein **11** IV 553.] It is best kept as the *disodium salt* [831-59-4] which decomposes on heating [Beilstein **11** H 199, **11** I 48, **11** II 213, **11** III 453.]

***m*-Benzenedisulfonyl chloride** [585-47-7] **M 275.1, m 63°**. Crystallise it from  $\text{CHCl}_3$  (EtOH free, by passing through an alumina column) or  $^*\text{C}_6\text{H}_6$ /petroleum ether and dry it at 20mm pressure. [Beilstein 11 IV 553.]

**Benzene-1,2-dithiol** [17534-15-5] **M 142.2, m 24-25°, 27-28°, b 110-112°,  $\text{pK}_{\text{Est}(1)} \sim 6.0$ ,  $\text{pK}_{\text{Est}(2)} \sim 9.4$** . Likely impurities are the oxidation products, the disulfides which could be polymeric. Dissolve it in aqueous NaOH until the solution is alkaline. Extract with  $\text{Et}_2\text{O}$  and discard the extract. Acidify with cold HCl (diluted 1:1 by volume with  $\text{H}_2\text{O}$ ) to Congo Red paper under  $\text{N}_2$  and extract it three times with  $\text{Et}_2\text{O}$ . Dry the  $\text{Et}_2\text{O}$  with  $\text{Na}_2\text{SO}_4$ , filter, evaporate and distil the residue under reduced pressure in an atmosphere of  $\text{N}_2$ . The distillate solidifies on cooling. [UV: Dewar et al. *J Chem Soc* 3076 1958, Grunwald & Berkowitz *J Am Chem Soc* 81 4939 1951, Ferretti *Org Synth Coll Vol V* 419 1973, Beilstein 6 IV 5651.]

**Benzenesulfinic acid** [618-41-7] **M 142.2, m 84°,  $\text{pK}^{25} 2.16(2.74)$** . The acid is purified by dissolving the Na salt in  $\text{H}_2\text{O}$ , acidifying to Congo Red paper with HCl and adding a concentrated solution of  $\text{FeCl}_3$  whereby Fe sulfinate precipitates. Collect the salt, wash it with a little  $\text{H}_2\text{O}$ , drain, suspend it in  $\text{H}_2\text{O}$  and add a slight excess of 1.5M aqueous NaOH. The  $\text{Fe}(\text{OH})_3$  precipitates, it is filtered off, the sulfinic acid in the aqueous solution is extracted with  $\text{Et}_2\text{O}$ , the extract is dried ( $\text{Na}_2\text{SO}_4$ ) and evaporated to give colourless crystals of benzenesulfinic acid **m 84°** which are stored under  $\text{N}_2$  in the dark, as it slowly oxidises in air to the sulfonic acid. [Beilstein 11 H 2, 11 I 3, 11 II 3, 11 III 3, 11 IV 3.]

**Benzenesulfonic acid** [98-11-3] **M 158.2, m 43-44°, 50-55°(anhydrous), 65-66°,  $\text{pK}^{25} -2.7, 0.70, (2.53?)$**  Purify benzenesulfonic acid by dissolving it in a small volume of distilled  $\text{H}_2\text{O}$  and stirring with slightly less than the theoretical amount of  $\text{BaCO}_3$ . When effervescence is complete and the solution is still acidic, filter off the insoluble barium benzenesulfonate. The salt is collected and dried to constant weight *in vacuo*, then suspended in  $\text{H}_2\text{O}$  and stirred with a little less than the equivalent (half mol.) of sulfuric acid. The insoluble  $\text{BaSO}_4$  (containing a little barium benzenesulfonate) is filtered off and the filtrate containing the free acid is evaporated in a high vacuum. The oily residue will eventually crystallise when completely anhydrous. A 32% commercial acid is allowed to fractionally crystallise at room temperature over  $\text{P}_2\text{O}_5$  in a vacuum desiccator giving finally colourless deliquescent plates **m 52.5°**. The anhydrous crystalline acid is deliquescent and should be stored over anhydrous  $\text{Na}_2\text{SO}_4$  in the dark, and should be used in subdued sunlight as it darkens under sunlight. The main impurity is Fe which readily separates as the Fe salt in the early fractions [Taylor & Vincent *J Chem Soc* 3218 1952]. The *S*-benzylisothiuronium salt has **m 148°** (from EtOH/ $\text{H}_2\text{O}$ ). It is a strong acid, and an IRRITANT to the skin and eyes. [See Adams & Marvel *Org Synth Coll Vol I* 84 1941, Michael & Adair *Chem Ber* 10 585 1877, Beilstein 11 IV 27.]

**Benzenesulfonic anhydride** [512-35-6] **M 298.3, m 88-91°**. Crystallise the anhydride from  $\text{Et}_2\text{O}$  (**m 88.5-91.5°**),  $\text{CHCl}_3$  or chlorobenzene (**m 90-92°**). Store it dry. [Field *J Am Chem Soc* 74 394 1952, Beilstein 11 H 3, 11 I 11, 11 II 23, 11 III 50, 11 IV 50.]

**Benzenesulfonyl chloride** [98-09-9] **M 176.6, m 14.5°, b 120°/10mm, 251.2°/760mm(dec),  $d_4^{20} 1.384$** . Distil the sulfonyl chloride, preferably under a vacuum, then treat it with 3mole % each of toluene and  $\text{AlCl}_3$ , and allow it to stand overnight. The sulfonyl chloride is distilled off at 1mm pressure and then carefully fractionally distilled at 10mm in an all-glass column. [Adams & Marvel *Org Synth Coll Vol I* 84 1941, Jensen & Brown *J Am Chem Soc* 80 4042 1958, Beilstein 11 IV 49.] It is **TOXIC**.

**Benzene-1,2,4,5-tetracarboxylic (pyromellitic acid)** [89-05-4] **M 254.2, m 276°, 281-284°,  $\text{pK}_1^{25} 1.87, \text{pK}_2^{25} 2.72, \text{pK}_3^{25} 4.30, \text{pK}_4^{25} 5.52$** . Dissolve it in 5.7 parts of hot dimethylformamide, decolorise, filter and cool. The precipitate is collected and dried in air. The solvent is removed by heating in an oven at 150-170° for several hours. It also crystallises from water. [Beilstein 9 H 997, 9 IV 3804.]

**Benzene-1,2,4,5-tetracarboxylic dianhydride (pyromellitic dianhydride)** [89-32-7] **M 218.1, m 286°, b 397-400°/760mm**. Crystallise the dianhydride from ethyl methyl ketone or dioxane. Dry, and sublime it *in vacuo*. [Beilstein 19 H 196, 19/5 V 407.]

**Benzene-1,2,3-tricarboxylic (hemimellitic) acid ( $\text{H}_2\text{O}$ )** [36362-97-7] **M 210.1, m 190°(dec),  $\text{pK}_1^{25} 2.62, \text{pK}_2^{25}$**

**3.82, pK<sub>3</sub><sup>25</sup> 5.51.** Crystallise the acid from water. [*Beilstein* 9 H 976, 9 IV 3745.]

**Benzene-1,3,5-tricarboxylic (trimesic or trimellitic) acid** [554-95-0] M 210.1, m 360°(dec), pK<sub>1</sub><sup>25</sup> 2.64, pK<sub>2</sub><sup>25</sup> 3.71, pK<sub>3</sub><sup>25</sup> 5.01. Crystallise the acid from water. The *trimethyl ester* has m 144° (from MeOH or MeOH/H<sub>2</sub>O). [*Beilstein* 9 H 978, 9 IV 3747.]

**1,2,4-Benzenetriol** [533-73-3] M 126.1, m 140.5-141°(sintering at 139°), pK<sub>1</sub><sup>20</sup> 9.08, pK<sub>2</sub><sup>20</sup> 11.82. Crystallise the triol from Et<sub>2</sub>O or Et<sub>2</sub>O/EtOH, and dry it in a vacuum. The *picrate* forms orange-red needles m 96°. [*Beilstein* 6 H 1087, 6 I 541, 6 II 1071, 6 III 6276.]

**Benzethonium chloride (Hyamine 1622, [diisobutylphenoxyethoxyethyl]dimethylbenzyl-ammonium chloride, (N,N-dimethyl-N-[2-[2-[4-(1,1,3,3-tetramethylbutyl)phenoxy]-ethoxy]-ethyl]-benzenemethan-ammonium chloride)** [121-54-0] M 448.1, m 164-166° (sinters at 120°, monohydrate). Crystallise it from boiling acetone after filtering or from CHCl<sub>3</sub>/petroleum ether. The precipitate is filtered off, washed with diethyl ether and dried for 24 hours in a vacuum desiccator. It is a cationic antiseptic surfactant which forms crystals also from a 1:9 MeOH/Et<sub>2</sub>O mixture. It foams in water. [*Beilstein* 12 IV 2187.]

**Benzhydrol (diphenylmethanol)** [91-01-0] M 184.2, m 69°, b 297°/748mm, 180°/20mm. Crystallise benzhydrol from hot H<sub>2</sub>O or petroleum ether (b 60-70°), petroleum ether containing a little \*benzene, from CCl<sub>4</sub>, or EtOH (1ml/g). An additional purification step includes passage of a \*benzene solution through an activated alumina column. It sublimes in a vacuum. Also recrystallise it three times from MeOH/H<sub>2</sub>O [*Naguib J Am Chem Soc* 108 128 1986]. [*Beilstein* 6 IV 4648.]

§ A commercial polystyrene supported version is available.

**Benzidine (4,4'-diaminobiphenyl)** [92-87-5] M 184.2, m 128-129°, pK<sub>1</sub><sup>20</sup> 3.85, pK<sub>2</sub><sup>20</sup> 4.95. Its solution in \*benzene is decolorised by percolating through two 2-cm columns of activated alumina, then concentrated until benzidine crystallises on cooling. Recrystallise alternately from EtOH and \*benzene to constant absorption spectrum [*Carlin et al. J Am Chem Soc* 73 1002 1951]. It has also been crystallised from hot water (charcoal) and from diethyl ether. Dry it under vacuum in an Abderhalden pistol. Store it in the dark in a stoppered container. **CARCINOGENIC.** [*Beilstein* 13 IV 364.]

**Benzidine dihydrochloride** [531-85-1] M 257.2, m >250°(dec). Crystallise the salt by dissolving in hot H<sub>2</sub>O, and adding conc HCl to the slightly cooled solution. **CARCINOGENIC.** [*Beilstein* 13 IV 365.]

**Benzil** [134-81-6] M 210.2, m 96-96.5°. Crystallise benzil from \*benzene after washing with alkali. (Crystallisation from EtOH did not free benzil from material reacting with alkali.) [*Hine & Howarth J Am Chem Soc* 80 2274 1958.] It has also been crystallised from CCl<sub>4</sub>, diethyl ether or EtOH [*Inoue et al. J Chem Soc, Faraday Trans 1* 82 523 1986]. [*Beilstein* 7 IV 2502.]

**Benzilic acid (diphenylglycollic acid)** [76-93-7] M 228.3, m 150°, pK<sup>18</sup> 3.06. Crystallise benzilic acid from \*benzene (ca 6ml/g), or hot H<sub>2</sub>O. [*Beilstein* 10 IV 1256.]

**Benzil monohydrazone** [5433-88-7] M 224.3, m 151°(dec). Crystallise it from EtOH. The *monoacetyl hydrazone* has m 91° (from EtOH). [*Metze & Meyer Chem Ber* 90 483 1959, *Beilstein* 7 I 394.]

**α-Benzil monoxime** [14090-77-8], [*E* 574-15-2], [*Z* 574-16-3] M 105.1, m 140°. The *trans-isomer* crystallises from \*C<sub>6</sub>H<sub>6</sub> (must not use animal charcoal) and has m 140°. The *cis-isomer* also crystallises from \*C<sub>6</sub>H<sub>6</sub> but crystals have 0.5\*C<sub>6</sub>H<sub>6</sub> (m 62-63°), and the solvent free compound has m 112°. [*Beilstein* 7 III 3812, 7 IV 2504.]

**Benzo[*a*]biphenylene** [252-47-1] M 202.2, m 72-73° (compare with β-isomer). It forms yellow needles from MeOH and sublimes *in vacuo* (m 72.0-72.8°). The 2,4,7-trinitrofluorenone complex crystallises as black needles m 201.5-202.5°. [*Cava & Stucker J Am Chem Soc* 77 6022 1955, *Barton et al. J Chem Soc(C)* 1276

1967, *Beilstein* 5 IV 2462.]

**Benzo[*b*]biphenylene** [259-56-3] **M 202.2, 242.6-243.6°**. It forms yellow crystals from \*C<sub>6</sub>H<sub>6</sub>/cyclohexane **m** 234-245° (sublimation). The 2,4,7-trinitrofluorenone complex crystallises as red needles from \*C<sub>6</sub>H<sub>6</sub>/MeOH **m** 214-216°. It has been sublimed *in vacuo*. [Jensen & Coleman *Tetrahedron Lett* No 20 7 1959, Barton et al. *J Chem Soc Perkin Trans 1* 1967 1986, *Beilstein* 5 IV 2462.]

**Benzoic acid** [65-85-0] **M 122.1, m 122.6-123.1°, pK<sup>25</sup> 4.12**. For use as a volumetric standard, analytical reagent grade benzoic acid should be carefully fused to *ca* 130° (to dry it) in a platinum crucible, and then powdered in an agate mortar. Benzoic acid has been crystallised from boiling water (charcoal), aqueous acetic acid, glacial acetic acid, \*C<sub>6</sub>H<sub>6</sub>, aqueous EtOH, petroleum ether (b 60-80°), and from EtOH solution by adding water. It is readily purified by fractional crystallisation from its melt and by sublimation in a vacuum at 80°. The *S*-benzylisothiuronium salt has **m** 167° (from EtOH/H<sub>2</sub>O). [*Beilstein* 9 IV 273.]

**Benzoic anhydride** [93-97-0] **M 226.2, m 42°**. Free it from benzoic acid by washing with NaHCO<sub>3</sub>, then water, and drying. Crystallise it from \*benzene (0.5ml/g) by adding just enough petroleum ether (b 40-60°) to cause cloudiness, then cool in ice. It can be distilled at 210-220°/20mm. [*Beilstein* 19 IV 550.]

**(±)-Benzoin (2-hydroxy-2-phenylacetophenone)** [119-53-9] **M 212.3, m 137°**. Crystallise benzoin from CCl<sub>4</sub>, hot EtOH (8ml/g), or 50% acetic acid. Also crystallise it from high purity \*benzene, then twice from high purity MeOH, to remove fluorescent impurities [Elliott & Radley *Anal Chem* 33 1623 1961]. It can be sublimed. [*Beilstein* 8 IV 1279.]

**(±)-α-Benzoinoxime** [441-38-3] **M 227.3, m 151°**. Crystallise the oxime from diethyl ether. It is used for the spectroscopic determination of Cu<sup>2+</sup>, Pd<sup>2+</sup>, Pt<sup>4+</sup>, Rh<sup>3+</sup> and V<sup>5+</sup> [Singh et al. *Talanta* 26 425 1979, *Beilstein* 8 IV 1282.]

**Benzonitrile** [100-47-0] **M 103.1, f -12.9°, b 191.1°, d<sub>4</sub><sup>20</sup> 1.010, n<sub>D</sub><sup>20</sup> 1.528**. Dry benzonitrile with CaSO<sub>4</sub>, CaCl<sub>2</sub>, MgSO<sub>4</sub> or K<sub>2</sub>CO<sub>3</sub>, and distil it from P<sub>2</sub>O<sub>5</sub> in an all-glass apparatus, under reduced pressure (**b** 69°/10mm), collecting the middle fraction. Distillation from CaH<sub>2</sub> causes some decomposition of benzonitrile. Isonitriles can be removed by preliminary treatment with conc HCl until the odour of isonitrile (carbylamine) has gone, followed by preliminary drying with K<sub>2</sub>CO<sub>3</sub>. (This treatment also removes amines.) Steam distil (to remove small quantities of carbylamine). The distillate is extracted into ether, washed with dilute Na<sub>2</sub>CO<sub>3</sub>, dried overnight with CaCl<sub>2</sub>, and the ether is removed by evaporation. The residue is distilled at 40mm (**b** 96°) [Kice et al. *J Am Chem Soc* 82 834 1960].

Conductivity grade benzonitrile (specific conductance 2 x 10<sup>-8</sup> mho) is obtained by treatment with anhydrous AlCl<sub>3</sub>, followed by rapid distillation at 40-50° under vacuum. After washing with alkali and drying with CaCl<sub>2</sub>, the distillate is redistilled in a vacuum several times at 35° before fractionally crystallising several times by partial freezing. It is dried over finely divided activated alumina from which it is withdrawn when required [Van Dyke & Harrison *J Am Chem Soc* 73 402 1951]. [*Beilstein* 9 IV 892.]

**Benzo[*ghi*]perylene (1,12-benzoperylene)** [191-24-2] **M 276.3, m 273°, 277-278.5°, 278-280°**. It forms light green crystals on recrystallisation from \*C<sub>6</sub>H<sub>6</sub> or xylene and sublimes at 320-340°/0.05mm [UV: Hopff & Schweizer *Helv Chim Acta* 42 2315 1959, Clar *Chem Ber* 65 846 1932, Fluoresc. Spectrum: Bowen & Brocklehurst *J Chem Soc* 3875 1954]. It also recrystallises from propan-1-ol [Altman & Ginsburg *J Chem Soc* 466 1959]. The 1,3,5-trinitrobenzene complex has **m** 310-313° (deep red crystals from \*C<sub>6</sub>H<sub>6</sub>), the picrate has **m** 267-270° (dark red crystals from \*C<sub>6</sub>H<sub>6</sub>), and the styphnate (2,4,6-trinitroresorcinol complex) has **m** 234° (wine red crystals from \*C<sub>6</sub>H<sub>6</sub>). [*Beilstein* 5 IV 2766.]

**3,4-Benzophenanthrene (benzo[*c*]phenanthrene)** [195-19-7] **M 228.3, m 68°**. Crystallise benzo[*c*]phenanthrene from EtOH, petroleum ether, or EtOH/Me<sub>2</sub>CO. [*Beilstein* 5 III 2378, 5 IV 2552.]

**Benzophenone** [119-61-9] **M 182.2, m 48.5-49°, pK<sup>25</sup> -6.0(-8.4) (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise it from MeOH, EtOH, cyclohexane, \*benzene or petroleum ether, then dry in a current of warm air and store it over BaO

or  $P_2O_5$ . It is also purified by zone melting and by sublimation [Itoh *J Phys Chem* **89** 3949 1985, Naguib et al. *J Am Chem Soc* **108** 128 1986, Gorman & Rodgers *J Am Chem Soc* **108** 5074 1986, Ohamoto & Teranishi *J Am Chem Soc* **108** 6378 1986, Naguib et al. *J Phys Chem* **91** 3033 1987]. [Beilstein 7 III 2048, 7 IV 1357.]

**Benzophenone oxime** [574-66-3] **M 197.2, m 142°, 143-144°, pK<sup>25</sup> 11.18**. Crystallise the oxime from MeOH (4ml/g). [Beilstein 7 II 355, 7 III 2063 1370.]

**Benzophenone-3,3',4,4'-tetracarboxylic dianhydride (BTDA)** [2421-28-5] **M 322.2, m 218.5-219.5°, 225.5°**. The main impurity is the free acid formed by hydrolysis (check for OH bands in the IR). This can be converted to the dianhydride by treating with  $Ac_2O$  (molar ratio of 4 to 1 of acid), heating at 110-120° for 1.5 to 2 hours, cooling to 0—5° and collecting the dianhydride. This is then dissolved in hot dioxane or  $Me_2CO$ , filtered and cooled to 10—15°. The moisture sensitive solid is collected and dried at 120—130° *in vacuo*. It has been sublimed at high vacuum. [Faberov et al. *J Org Chem USSR* **4** 153 (English translation) 1968.]

**Benzopinacol** [464-72-2] **M 366.5, m 170-180° (depends on heating rate), 171-173°**. Crystallise benzopinacol from EtOH. [Beilstein 6 IV 7053.]

**Benzo[a]pyrene (3,4-benzopyrene, benzo[def]chrysene)** [50-32-8] **M 252.3, m 177.5-178°, 179.0-179.5°**. A solution of 250mg of benzo[a]pyrene in 100ml of \*benzene is diluted with an equal volume of hexane, then passed through a column of alumina,  $Ca(OH)_2$  and Celite (3:1:1). The adsorbed material is developed with a 2:3 \*benzene/hexane mixture. (It showed as an intensely fluorescent zone.) The main zone is eluted with 3:1 acetone/EtOH, and is transferred into 1:1 \*benzene-hexane by adding  $H_2O$ . The solution is washed, dried with  $Na_2SO_4$ , evaporated and crystallised from \*benzene by the addition of MeOH [Lijinsky & Zechmeister *J Am Chem Soc* **75** 5495 1953]. Alternatively, it can be chromatographed on activated alumina, eluted with a cyclohexane/\*benzene mixture containing up to 8% \*benzene, and the solvent is evaporated under reduced pressure [Cahnmann *Anal Chem* **27** 1235 1955], and crystallised from EtOH [Nithipatikom & McGown *Anal Chem* **58** 3145 1986]. [Beilstein 5 III 2517, 5 IV 2687.] **CARCINOGENIC**.

**Benzo[e]pyrene (1,2-benzopyrene)** [192-97-2] **M 252.3, m 178-179°, 178-180°**. Purify it by passage through an  $Al_2O_3$  column (Woelm, basic, activity I) and elute with \* $C_6H_6$  and recrystallise from 2 volumes of EtOH/\* $C_6H_6$  (4:1). It forms colourless or light yellow prisms or needles. [Campbell *J Chem Soc* 3659 1954, Buchta & Kröger *Justus Liebigs Ann Chem* **705** 190 1967.] The 1,3,5-trinitrobenzene complex has **m 253-254°** (orange needles from EtOH), the picrate prepared by mixing 20mg in 1ml of \* $C_6H_6$  with 20mg of picric acid in 2ml \* $C_6H_6$ , collecting the deep red crystals, and recrystallising from \* $C_6H_6$  has **m 228-229°** [NMR: Cobb & Memory *J Chem Phys* **47** 2020 1967]. [Beilstein 5 III 2520, 5 IV 2689.] **CARCINOGEN**.

**p-Benzoquinone** [106-51-4] **M 108.1, m 115.7°**. Purify *p*-benzoquinone in one or more of the following ways: steam distillation followed by filtration and drying (e.g. in a desiccator over  $CaCl_2$ ), crystallisation from petroleum ether (b 80-100°), \*benzene (with, then without, charcoal), water or 95% EtOH, sublimation under vacuum (e.g. from room temperature to liquid  $N_2$ ). It slowly decomposes and should be stored, refrigerated, in an evacuated or sealed glass vessel in the dark. It should be resublimed before use. [Wolfenden et al. *J Am Chem Soc* **109** 463 1987, Beilstein 7 IV 2065.]

**1-Benzosuberone (6,7,8,9-tetrahydrobenzocyclohepten-5-one)** [826-73-3] **M 160.2, b 80-85°/0.5mm, 90-93°/1mm, 138-139°/12mm, 154°/15mm, 175-175°/40mm, d<sub>4</sub><sup>20</sup> 1.086, n<sub>D</sub><sup>20</sup> 1.5638**. Purify it by dissolving in toluene, washing with aqueous 5% NaOH, then brine, drying ( $MgSO_4$ ), and distilling. The 2,4-dinitrophenylhydrazone has **m 210.5°, 207-208°** (from  $CHCl_3/MeOH$ ). The *Z-O-picryloxime* has **m 156-157°** (from  $Me_2CO/MeOH$ ), the *E-O-picryloxime* has **m 107°**. The oxime has **m 106.5-107.5°**. [UV: Gilmore & Horton *J Am Chem Soc* **73** 1411 1951, Hedden & Brown *J Am Chem Soc* **75** 3744 1953, Huisgen et al. *Chem Ber* **90** 1844 1957, Beilstein 7 IV 1029.]

**Benzoylacetone (1-phenyl-1,3-butanedione)** [93-91-4] **M 162.2, m 58.5-59.0°**. Crystallise benzoylacetone from  $Et_2O$  or MeOH and dry it under vacuum at 40°. [Beilstein 7 IV 2151.]

**2-Benzoylbenzoic acid** [85-52-9] **M 226.2, m 126-129°, 129.2, 130°, pK<sup>25</sup> 3.54.** Recrystallise the acid from \*C<sub>6</sub>H<sub>6</sub> or cyclohexane, but it is best recrystallised by dissolving in a small volume of hot toluene and then adding just enough petroleum ether to cause turbidity, and cool. Dry it in a low vacuum at 80°. It can be sublimed at 230-240°/0.3mm [Bray et al. *J Chem Soc* 265 1957]. The *S*-benzylisothiuronium salt has **m 177-178°** (from EtOH). [Lewenz & Serijan *J Am Chem Soc* 75 4087 1953, *Beilstein* 10 H 747, 10 IV 2977.]

**3-Benzoylbenzoic acid** [579-18-0] **M 226.2, m 164-166°, 165°, pK<sub>Est</sub>~3.5.** Crystallise the acid from EtOH and sublime it at 160°/1mm. [*Beilstein* 10 H 752, 10 III 3304, 10 IV 2982.]

**4-Benzoylbenzoic acid** [611-95-0] **M 226.2, m 196.5-198°, 197-200°, pK<sub>Est</sub>~3.7.** Dissolve the acid in hot H<sub>2</sub>O by adding enough aqueous KOH solution till distinctly alkaline, filter and then acidify with drops of conc HCl. Filter off, wash the solid with cold H<sub>2</sub>O, dry it at 100°, and recrystallise it from EtOH. [Wertheim *J Am Chem Soc* 55 2540 1933, *Beilstein* 10 H 753, 10 IV 3305.]

**Benzoyl chloride** [98-88-4] **M 140.6, b 56°/4mm, 196.8°/745mm, d<sub>4</sub><sup>20</sup> 1.2120, n<sub>D</sub><sup>10</sup> 1.5537.** A solution of benzoyl chloride (300ml) in \*C<sub>6</sub>H<sub>6</sub> (200ml) is washed with two 100ml portions of cold 5% NaHCO<sub>3</sub> solution, separated, dried with CaCl<sub>2</sub> and distilled [Oakwood & Weisgerber *Org Synth* III 113 1955]. Repeated fractional distillation at 4mm Hg through a glass helices-packed column (avoiding porous porcelain or silicon-carbide boiling chips, and hydrocarbon or silicon greases on the ground joints) gave benzoyl chloride that did not darken on addition of AlCl<sub>3</sub>. Further purification is achieved by adding 3 mole% each of AlCl<sub>3</sub> and toluene, standing overnight, and distilling off the benzoyl chloride at 1-2mm [Brown & Jenzen *J Am Chem Soc* 80 2291 1958]. Refluxing for 2 hours with an equal weight of thionyl chloride before distillation has also been used. [*Beilstein* 9 IV 721.] **Strong IRRITANT.** Use in a fume cupboard.

**Benzoyl disulfide (dibenzoyl disulfide)** [644-32-6] **M 174.4, m 131.2-132.3°.** About 300ml of solvent is blown off from a filtered solution of dibenzoyl disulfide (25g) in acetone (350ml). The remaining acetone is decanted from the solid which is recrystallised first from 300ml of 1:1 (v/v) EtOH/ethyl acetate, then from 300ml of EtOH, and finally from 240ml of 1:1 (v/v) EtOH/ethyl acetate. The yield is about 40% [Pryor & Pickering *J Am Chem Soc* 84 2705 1962]. [*Beilstein* 9 H 424, 9 II 289, 9 III 1977.] *Handle in a fume cupboard because of TOXICITY and obnoxious odour.*

**Benzoylformic acid (phenylglyoxylic acid)** [611-73-4] **M 150.14, m 62-65°, 64.5-65.5°, 67°, b 84°/0.1mm, 163-167°/15mm, pK<sup>25</sup> 1.39 (1.79).** If the sample is oily, then it may contain H<sub>2</sub>O. In this case dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or KOH until crisp. For further purification dissolve 5.5g in hot CCl<sub>4</sub> (750ml), add charcoal (2g, this is necessary otherwise the acid may separate as an oil), filter, cool in ice-water until crystallisation is complete. Filter the acid off, and the solvent on the crystals is removed by keeping the acid (4.5g) in a vacuum desiccator for 2 days. Slightly yellow crystals are obtained. It can be recrystallised also from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, and can be distilled in a vacuum. The acid is estimated by titration with standard NaOH. The *phenylhydrazone* is recrystallised from EtOH, **m 163-164°**, and the *semicarbazone acid* has **m 259°(dec)** (from EtOH). The *methyl ester* distils at **137°/14mm, 110-111°/2mm, n<sub>D</sub><sup>20</sup> 1.5850.** [Baert & Kates *J Am Chem Soc* 67 1482 1945, Schaffer & Corey *J Org Chem* 24 1825 1959, *Beilstein* 10 H 654, 10 IV 2737.]

**Benzoyl isothiocyanate** [532-55-8] **M 163.2, m 25.5-26°, b 72.5-73°/6mm, 88-91°/20mm, 94-96°/21mm, 202.5-204°/724mm, 250-255°/atm, d<sub>4</sub><sup>20</sup> 1.213, n<sub>D</sub><sup>20</sup> 1.637.** Distil the isothiocyanate over a small amount of P<sub>2</sub>O<sub>5</sub>, whereby the distillate crystallises in prisms. It is readily hydrolysed by H<sub>2</sub>O to give benzamide and benzoylurea, but with NH<sub>3</sub> it gives *benzoylurea* **m 210°** which can be recrystallised from EtOH. [Hill & Degnan *J Am Chem Soc* 62 1595 1940, Terss & McEwen *J Am Chem Soc* 76 580 1954, Frank & Smith *Org Synth Coll Vol* III 735 1955, *Beilstein* 9 IV 777.]

**Benzoyl peroxide** [94-36-0] **M 242.2, m 95°(dec).** Dissolve benzoyl peroxide in CHCl<sub>3</sub> at room temperature and precipitate it by adding an equal volume of MeOH or petroleum ether. Similarly, it is precipitated from acetone by adding two volumes of distilled water. It has also been crystallised from 50% MeOH and from diethyl ether. Dry it under vacuum at room temperature for 24 hours. Store it in a desiccator in the dark at 0°. When purifying in the absence of water it can be **EXPLOSIVE**, and operations should be done on a very small

scale with adequate protection. Large amounts should be kept moist with water and stored in a refrigerator. [Kim et al. *J Org Chem* **52** 3691 1987, *Beilstein* **9** IV 777.]

***p*-Benzoylphenol (4-hydroxybenzophenone)** [1137-42-4] **M 198.2, m 133.4-134.8<sup>o</sup>, pK<sup>25</sup> 7.95.** Dissolve *p*-benzoylphenol in hot EtOH (charcoal), filter and cool. Alternatively, crystallise it once from EtOH/H<sub>2</sub>O and twice from \*benzene [Grunwald *J Am Chem Soc* **73** 4934 1951, Dryland & Sheppard *J Chem Soc Perkin Trans I* 125 1986]. [*Beilstein* **8** IV 1263.]

***N*-Benzoyl-*N*-phenylhydroxylamine** [304-88-1] **M 213.2, m 121-122<sup>o</sup>.** Recrystallise it from hot water, \*benzene Et<sub>2</sub>O/hexane or acetic acid. It complexes with metals. [*Beilstein* **15** III 8, **15** IV 7.]

***N*-Benzoyl-*o*-tolylhydroxylamine** [1143-74-4] **M 227.3, m 104<sup>o</sup>.** Recrystallise the hydroxylamine from aqueous EtOH. [*Beilstein* **15** III 8, **15** IV 7.]

**Benzyl acetate** [140-11-4] **M 150.2, m -51<sup>o</sup>, b 92-93<sup>o</sup>/10mm, 134<sup>o</sup>/102mm, 214.9<sup>o</sup>/760mm, d<sub>4</sub><sup>20</sup> 1.0562, n<sub>D</sub><sup>25</sup> 1.4994.** Purify the acetate by fractional distillation, preferably in a good vacuum. Values of n<sub>D</sub><sup>25</sup> of 1.5232-1.5242 are too high and should be nearer to 1.4994. [Merker & Scott *J Org Chem* **26** 5180 1961, *Beilstein* **6** IV 2262.]

**Benzyl acetoacetate** [5396-89-4] **M 192.2, b 130<sup>o</sup>/2mm, 156-157<sup>o</sup>/10mm, 162-167<sup>o</sup>/15mm, 275-277<sup>o</sup>/atm, d<sub>4</sub><sup>20</sup> 1.114, n<sub>D</sub><sup>20</sup> 1.514.** Fractionate the ester and collect fractions with the expected physical properties. Otherwise add *ca* 10% by weight of benzyl alcohol and heat in an oil bath (160-170<sup>o</sup>, open vessel) for 30 minutes during which time excess of benzyl alcohol will have distilled off, then fractionate. [Baker et al. *J Org Chem* **17** 77 1952, *Beilstein* **6** IV 2480.]

**4'-Benzylacetophenone** [782-92-3] **M 210.3, m 37<sup>o</sup>, 38<sup>o</sup>, 39<sup>o</sup>, b 197-198<sup>o</sup>/9mm, 209-210<sup>o</sup>/15mm.** Distil it in a vacuum, then recrystallise it from EtOH (*ca* 1ml/g). The *oxime* has **m 99.5<sup>o</sup>** (from 60% aqueous EtOH). [*Beilstein* **7** H 449, **7** III 2176.]

**Benzyl alcohol** [100-51-6] **M 108.1, f -15.3<sup>o</sup>, b 205.5<sup>o</sup>, 93<sup>o</sup>/10mm, 205.5<sup>o</sup>/atm, d<sub>4</sub><sup>20</sup> 0.981, n<sub>D</sub><sup>20</sup> 1.54033, pK<sup>25</sup> 15.4.** It is usually purified by careful fractional distillation under reduced pressure in the absence of air. Benzaldehyde, if present, can be detected by UV absorption at 283nm. It has also been purified by shaking with aqueous KOH and extracting with peroxide-free diethyl ether. After washing with water, the extract is treated with saturated NaHS solution, filtered, washed, dried with CaO and distilled under reduced pressure [Mathews *J Am Chem Soc* **48** 562 1926]. Peroxy compounds can be removed by shaking with a solution of Fe<sup>2+</sup> followed by washing the alcohol layer with distilled water and fractionally distilling it. [*Beilstein* **6** IV 2222.]

**Benzylamine** [100-46-9] **M 107.2, b 178<sup>o</sup>/742mm, 185<sup>o</sup>/768mm, d<sub>4</sub><sup>20</sup> 0.981, n<sub>D</sub><sup>20</sup> 1.5392, pK<sup>25</sup> 9.33.** Dry it with NaOH or KOH, then distil it under N<sub>2</sub>, through a column packed with glass helices, taking the middle fraction. Also distil it from zinc dust under reduced pressure. The *picrate* has **m 196<sup>o</sup>** (from EtOH), and the *p*-toluenesulfonamide has **m 116<sup>o</sup>** (from MeOH). [*Beilstein* **12** IV 2155.]

**Benzylamine hydrochloride** [3287-99-8] **M 143.6, m 248<sup>o</sup> (rapid heating).** Crystallise the salt from water. [*Beilstein* **12** IV 2155.]

***N*-Benzylaniline (*N*-phenylbenzylamine)** [103-32-2] **M 183.4, m 36<sup>o</sup>, b 306-307<sup>o</sup>, d<sub>4</sub><sup>20</sup> 1.061, pK<sup>25</sup> 4.04.** Crystallise the amine from petroleum ether (b 60-80<sup>o</sup>) (*ca* 0.5ml/g). The *picrate* has **m 113<sup>o</sup>** (from Et<sub>2</sub>O). [*Beilstein* **12** H 1023, **12** I 449, **12** II 548, **12** III 2215, **12** IV 2172.]

**Benzyl bromide** [100-39-0] **M 171.0, m -4<sup>o</sup>, b 85<sup>o</sup>/12mm, 192<sup>o</sup>/760mm, d<sub>4</sub><sup>20</sup> 1.438, n<sub>D</sub><sup>20</sup> 1.575.** Wash benzyl chloride with conc H<sub>2</sub>SO<sub>4</sub> (CARE), water, 10% Na<sub>2</sub>CO<sub>3</sub> or NaHCO<sub>3</sub> solution, and again with water. Dry it with CaCl<sub>2</sub>, Na<sub>2</sub>CO<sub>3</sub> or MgSO<sub>4</sub> and fractionally distil it in the dark, under reduced pressure. It has also been thoroughly degassed at 10<sup>-6</sup> mm and redistilled in the dark. This gives material with λ<sub>max</sub> at (MeCN): 226nm (ε 8200) [Mohammed & Kosower *J Am Chem Soc* **93** 2709 1971]. [*Beilstein* **5** IV 829.] *Handle in a fume cup-*

board, extremely LACHRYMATORY.

**Benzyl bromoacetate** [5437-45-6] M 229.1, b 96-98°/0.1mm, 146°/12mm, 166-170°/22mm,  $d_4^{20}$  1.444,  $n_D^{25}$  1.5412. Dilute the ester with Et<sub>2</sub>O, wash it with 10% aqueous NaHCO<sub>3</sub>, H<sub>2</sub>O, dry (MgSO<sub>4</sub>) and fractionate it using a Fenske (glass helices packing) column. [Bergmann & Szinai *J Chem Soc* 1521 1956, *Beilstein* 6 IV 2265.] LACHRYMATORY.

**N-Benzyl-tert-butylamine** (*N*-tert-butylbenzylamine) [3378-72-1] M 163.3, b 91°/12mm, 109-110°/25mm, 218-220°/atm,  $d_4^{20}$  0.899,  $n_D^{25}$  1.4942, pK<sup>25</sup> 10.19. Dissolve the amine in Et<sub>2</sub>O, dry it over KOH pellets, filter and fractionate it in a N<sub>2</sub> atmosphere to avoid reaction with CO<sub>2</sub> from the air. The *hydrochloride* has m 245-246°(dec) (from MeOH/Me<sub>2</sub>CO) and the *perchlorate* has m 200-201°. [Freidfelder et al. *J Am Chem Soc* 80 4320 1958, *Beilstein* 12 IV 2166.]

**Benzyl carbamate** [621-84-1] M 151.2, m 86°, 86-88°, 90-91°. If it smells of NH<sub>3</sub>, then dry it in a vacuum desiccator and recrystallise it from 2 volumes of toluene and dry it in a vacuum desiccator again. It forms glistening plates from toluene, and can be recrystallised from H<sub>2</sub>O [Martell & Herbst *J Org Chem* 6 878 1941, Carter et al. *Org Synth Coll Vol III* 168 1955]. [*Beilstein* 6 IV 2278.]

**Benzyl chloride** [100-44-7] M 126.6, m 139°, b 63°/8mm,  $d_4^{20}$  1.100,  $n_4^{20}$  1.538. Dry it with MgSO<sub>4</sub> or CaSO<sub>4</sub>, or reflux it with fresh Ca turnings, then fractionally distil it under reduced pressure, collecting the middle fraction and storing it over CaH<sub>2</sub> or P<sub>2</sub>O<sub>5</sub>. It has also been purified by passage through a column of alumina. *Alternatively*, it is dried over MgSO<sub>4</sub> and distilled in a vacuum. The middle fraction is degassed by several freeze-thaw cycles and then fractionated in an 'isolated fractionating column' (which has been evacuated and sealed off at ~10<sup>-6</sup> mm) over a steam bath. The middle fraction is retained. The final samples are distilled in a vacuum from this sample and again retaining the middle fraction. The purity is >99.9% (no other peaks are visible by GLC, and the NMR spectrum is consistent with the structure. [Mohammed & Kosower *J Am Chem Soc* 93 1709 1971, *Beilstein* 5 IV 809.] IRRITANT and strongly LACHRYMATORY.

**N-Benzyl-β-chloropropionamide** [24752-66-7] M 197.7, m 94°, 96°. Crystallise the amide from MeOH. [12 III 2257, 12 IV 2234.]

**Benzyl cinnamate** [103-41-3] M 238.3, m 34-35°, 39°, b 154-157°/0.5mm, 228-230°/22mm. Recrystallise the ester to a constant melting point from 95% EtOH. It has the odour of balsam. *Alternatively*, dissolve it in Et<sub>2</sub>O, wash it with 10% aqueous Na<sub>2</sub>CO<sub>3</sub>, H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and fractionate it under reduced pressure using a short Vigreux column. It decomposes when boiled at atmospheric pressure. [Eliel & Anderson *J Am Chem Soc* 74 547 1952, Bender & Zerner *J Am Chem Soc* 84 2550 1962, *Beilstein* 9 IV 2012.]

**Benzyl cyanide** [140-29-4] M 117.1, b 100°/8mm, 233.5°/760mm,  $d_4^{20}$  1.015,  $n_4^{20}$  1.523. Any benzyl isocyanide impurity can be removed by shaking vigorously with an equal volume of 50% H<sub>2</sub>SO<sub>4</sub> at 60°, washing with saturated aqueous NaHCO<sub>3</sub>, then half-saturated NaCl solution, drying and fractionally distilling under reduced pressure. Distillation from CaH<sub>2</sub> causes some decomposition of this compound: it is better to use P<sub>2</sub>O<sub>5</sub>. Other purification procedures include passage through a column of highly activated alumina, and distillation from Raney nickel. *Precautions should be taken because of possible formation of free TOXIC cyanide, use an efficient fume cupboard.* [*Beilstein* 9 IV 1663.]

**N-Benzyl dimethylamine** [103-83-3] M 135.2, b 66-67°/15mm, 83-84°/30mm, 98-99°/24mm, 181°/760mm;  $d_4^{20}$  0.898,  $n_D^{20}$  1.516, pK<sup>25</sup> 8.91. Dry the amine over KOH pellets and fractionate it over Zn dust in a CO<sub>2</sub>-free atmosphere. It has a pK<sup>a25</sup> of 8.25 in 45% aqueous EtOH. Store it under N<sub>2</sub> or in a vacuum. The *picrate* has m 94-95°, and the *picrolonate* has m 151° (from EtOH). [Skita & Keil *Chem Ber* 63 34 1930, Coleman *J Am Chem Soc* 55 3001 1933, Devereux et al. *J Chem Soc* 2845 1957.] The *tetraphenyl borate salt* has m 182-185°. [Crane *Anal Chem* 28 1794 1956, *Beilstein* 12 IV 2161.]

**Benzyl dimethyloctadecylammonium chloride** [122-19-0] M 442.2, m 150-158° (sinters at 120°). Crystallise the salt from acetone, EtOAc or EtOAc/Et<sub>2</sub>O. [Sumiki et al. *J Agric Chem Soc Jpn* 26 325 1952,

*Chem Abstr* 3505 1953, *Beilstein* 12 III 2212, 12 IV 2168.]

**Benzyl ether (dibenzyl ether)** [103-50-4] M 198.3, b 298°, 158-160°/0.1mm,  $d_4^{20}$  1.043,  $n_D^{20}$  1.54057. Reflux the ether over sodium, then distil it under reduced pressure. It been purified by fractional freezing. [*Beilstein* 6 IV 2240.]

**N-Benzyl-N-ethylaniline** [92-59-1] M 221.3, b 212-222°/54mm, 285-286°/710mm, 312-313°/atm (dec),  $d_4^{20}$  1.029,  $n_D^{20}$  1.595,  $pK_{Est}$  ~4.6. Dry the amine over KOH pellets and fractionate it. The *picrate* crystallises from \*C<sub>6</sub>H<sub>6</sub> as lemon yellow crystals m 126-128° (softening at 120°). [Forrest et al. *J Chem Soc* 303 1951, IR: Hill & Meakins *J Chem Soc* 760 1958, *Beilstein* 12 H 1026, 12 IV 2176.]

**Benzyl ethyl ether** [539-30-0] M 136.2, b 186°, 65°/10mm,  $d_4^{20}$  0.949,  $n_D^{20}$  1.4955. Dry the ether with CaCl<sub>2</sub> or NaOH, then fractionally distil it. [Letzinger & Pollart *J Am Chem Soc* 78 6079 1956, *Beilstein* 6 III 1454, 6 IV 2229.]

**Benzyl ethyl ketone (1-phenylbutan-2-one)** [1007-32-5] M 148.2, b 49-49.5°/0.01mm, 66-69°/1mm, 83-85°/5mm, 101-102°/10mm, 229-233°/atm,  $d_4^{20}$  0.989,  $n_D^{25}$  1.5015. Purify the ketone by fractionation using an efficient column. It can be converted into the *oxime* which is distilled, b 117-118°/2mm, 145-146°/15mm,  $d_{25}^{25}$  1.036,  $n_D^{25}$  1.5363; decompose the oxime, and the ketone is redistilled. It can also be purified *via* the *semicarbazone* which has m 154-155°. [Meyers et al. *J Am Chem Soc* 77 5655 1955, Hass et al. *J Org Chem* 15 8 1950, *Beilstein* 7 IV 712.]

**O-Benzylhydroxylamine hydrochloride** [2687-43-6] M 159.6, m 234-238°(sublimes),  $pK_{Est}$  ~5.9. Recrystallise the hydrochloride from H<sub>2</sub>O or EtOH. [*Beilstein* 6 IV 2562.]

**N-Benzylideneaniline** [538-51-2] M 181.2, m 48° (54°), 56°, b 310°/760mm. It is steam volatile and crystallises from \*benzene or 85% EtOH. The *picrate* has m 159°. [*Beilstein* 12 H 195, 12 I 169, 12 II 113, 12 III 319, 12 IV 311.]

**Benzylidene malononitrile** [2700-22-3] M 154.2, m 83-84°, 87°. Recrystallise the nitrile from EtOH [Bernasconi et al. *J Am Chem Soc* 107 3612 1985]. It has  $\lambda_{max}$  at 307nm (EtOH). [*Beilstein* 9 H 895, 9 II 640, 9 III 4380, 9 IV 3462.]

**Benzyl isocyanate** [3173-56-6] M 133.2, b 82-84°/10mm, 87°/14mm, 95°/17mm, 101-104°/33mm,  $d_4^{20}$  1.08,  $n_D^{20}$  1.524. Purify the isocyanate by fractionation through a two-plate column. It is a viscous liquid and is TOXIC. [Howarth et al. *J Chem Soc* 182 1947, Ferstandig & Scherrer *J Am Chem Soc* 81 4838 1959, IR: Derkosch et al. *Monatsh Chem* 88 35 1957, *Beilstein* 12 IV 2276.]

**Benzyl isothiocyanate** [622-78-6] M 149.2, b 123-124°/1mm, 138-140°/20mm, 255-260°/atm,  $d_4^{20}$  1.1234,  $n_D^{20}$  1.6039. Dissolve benzyl isothiocyanate in Et<sub>2</sub>O, filter, if there is any solid, and distil it through an efficient column at 11mm with a bath temperature at ca 150°. Characterise it by reacting (0.5ml) in EtOH (1ml) with 50% NH<sub>2</sub>NH<sub>2</sub>·H<sub>2</sub>O (2 ml) to give 4-benzylthiosemicarbazide as colourless needles which are recrystallised from EtOH, m 130°. [Hoggarth & Young *J Chem Soc* 1582 1950, Schmidt et al. *Justus Liebigs Ann Chem* 612 11 1958, IR and UV: Svátek et al. *Acta Chem Scand* 13 442 1959, *Beilstein* 12 IV 2276.]

**S-Benzylisothiuronium chloride** [538-28-3] M 202.7, two forms, m 150° and 175°,  $pK_{Est}$  ~9.8 (free base). Crystallise the chloride from 0.2M HCl (2ml/g) or EtOH and it dry in air. [*Beilstein* 6 III 1600.]

**Benzylmalonic acid** [616-75-1] M 194.2, m 121°,  $pK_1^{25}$  2.91,  $pK_2^{25}$  5.87. Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub>. [*Beilstein* 9 IV 3357.]

**Benzyl mercaptan** [100-53-8] M 124.2, b 70.5-70.7°/9.5mm,  $d_4^{20}$  1.058,  $n_D^{20}$  1.5761,  $pK^{25}$  9.43. Purify benzyl mercaptan *via* the mercury salt [see Kern *J Am Chem Soc* 75 1865 1953], which crystallises from \*benzene as needles (m 121°), and then dissolve it in CHCl<sub>3</sub>. Pass H<sub>2</sub>S gas through the solution to regenerate

the mercaptan. The HgS that precipitates is filtered off and washed thoroughly with  $\text{CHCl}_3$ . The filtrate and washings are evaporated to remove  $\text{CHCl}_3$ ; then the residue is fractionally distilled under reduced pressure [Mackle & McClean, *Trans Faraday Soc* **58** 895 1962]. [*Beilstein* **6** IV 2632.]

**(-)-N-Benzyl-N-methylephedrinium bromide [benzyl(2-hydroxy-1-methyl-2-phenethyl) dimethylammonium bromide]** [58648-09-2] **M 350.3, m 209-211°, 212-214°,  $[\alpha]_{\text{D}}^{25}$  -3.8° (c 1.45, MeOH),  $[\alpha]_{\text{D}}^{20}$  -5.3° (c 1.45, MeOH)**. Recrystallise the bromide from MeOH/Et<sub>2</sub>O. [Horner & Brich *Justus Liebigs Ann Chem* 710 1978.] The *chloride* is recrystallised from EtOAc/*n*-hexane, **m 198-199°  $[\alpha]_{\text{D}}^{25}$  -8.67° (c 1.45, MeOH)**. [Julia et al. *J Chem Soc, Perkin Trans 1* 574 1981, *Beilstein* **13** IV 1890.]

**Benzyl 4-nitrophenyl carbonate** [13795-24-9] **M 273.2, m 78-80°**. Dissolve the carbonate in Et<sub>2</sub>O, wash with H<sub>2</sub>O (3x) and saturated aqueous NaCl, dry (MgSO<sub>4</sub>), evaporate this in a vacuum and recrystallise the residue from a small volume of MeOH, **m 78-79°**. Alternatively, dissolve it in Et<sub>2</sub>O, wash it with N HCl (2x), 0.5N NaHCO<sub>3</sub> (4x) then H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate the Et<sub>2</sub>O and recrystallise the residue from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, **m 79-80°**. The *2-nitro-isomer* has **m 27-28°, b 151°/11mm**. [Khosla et al. *Indian J Chem* **5** 279 1967, Wolman et al. *J Chem Soc (C)* 596 1976, *Beilstein* **6** IV 2277.]

**Benzyloxyacetyl chloride** [19810-31-2] **M 184.6, b 81°/0.2mm, 84-87°/0.4mm, 105-107°/5mm,  $d_4^{20}$  1.19,  $n_{\text{D}}^{20}$  1.523**. Check the IR to see if there are OH bands. If so, then it may be contaminated with free acid formed by hydrolysis. Add oxalyl chloride (amount depends on contamination and needs to be judged, *ca* 3mols), heat at 50° in the absence of moisture for 1 hour and fractionate twice, **b 81°/0.2mm** (with bath temperature at 81°). Excessive heating results in decomposition to give benzyl chloride. The *anilide* is formed by adding aniline in  $\text{CHCl}_3$  solution and has **m 49°**. [Fischer & Gohlke *Helv Chim Acta* **16** 1130 1933, *Beilstein* **6** IV 2470.]

**3-Benzyloxybenzoic acid** [69026-14-8] **M 228.2, m 133-137°, 135.5-136°,  $\text{pK}_{\text{Est}} \sim 4.1$** . Recrystallise the acid from acetic acid (**m 137-138°**) [Kipping & Wren *J Chem Soc* 3246 1957, *Beilstein* **10** III 247, **10** IV 316.]

**Benzyloxybutan-2-one** [6278-91-7] **M 178.2, b 90-92°/0.1mm, 88-91°/0.5mm, 121-126°/5mm,  $d_4^{20}$  1.0275,  $n_{\text{D}}^{20}$  1.5040**. Dissolve the ketone in  $\text{CHCl}_3$ , wash with H<sub>2</sub>O, aqueous saturated NaHCO<sub>3</sub>, H<sub>2</sub>O, dry (MgSO<sub>4</sub>), evaporate the  $\text{CHCl}_3$ , and fractionate it. [Hoffman et al. *J Am Chem Soc* **79** 2316 1957, *Beilstein* **6** IV 2255.]

**Benzyloxycarbonyl chloride (Cbz-Cl, BOC-Cl, benzyl chloroformate)** [501-53-1] **M 170.6, b 103°/20mm,  $d_4^{20}$  1.195,  $n_{\text{D}}^{20}$  1.5190**. The commercial material is usually better than 95% pure and may contain some toluene, benzyl alcohol, benzyl chloride and HCl. After long storage, e.g. two years at 4°, Greenstein and Winitz [*The Chemistry of the Amino Acids* **Vol 2** p. 890, J Wiley and Sons NY, 1961] recommended that the liquid should be flushed with a stream of dry air, filtered and stored over sodium sulfate to remove CO<sub>2</sub> and HCl which are formed by decomposition. It may further be distilled from an oil bath at a temperature below 85° because Thiel and Dent [*Annalen* **301** 257 1898] stated that benzyloxycarbonyl chloride decarboxylates to benzyl chloride slowly at 100° and vigorously at 155°. Redistillation at higher vacuum below 85° yields material which shows no other peaks than those of benzyloxycarbonyl chloride by NMR spectroscopy. [*Beilstein* **6** IV 2278.] **LACHRYMATORY and TOXIC.**

***p*-(Benzyloxy)phenol** [103-16-2] **M 200.2, m 122.5°,  $\text{pK}_{\text{Est}} \sim 10.1$** . Crystallise it from EtOH or H<sub>2</sub>O, and dry (P<sub>2</sub>O<sub>5</sub>) under vacuum. [Walter et al. *J Am Chem Soc* **108** 5210 1986, *Beilstein* **6** IV 5778.]

***S*-(*-*)-3-Benzyloxypropan-1,2-diol** [17325-85-8] **M 182.2, m 24-26°, b 117-118°/10<sup>-4</sup>mm, 115-116°/0.02mm, 121-123°/0.2mm,  $d_4^{20}$  1.1437,  $n_{\text{D}}^{22}$  1.5295,  $[\alpha]_{\text{D}}^{25}$  -5.9° (neat)**. Purify the diol by repeated fractional distillation. [Baer et al. *J Biol Chem* **230** 447 1958, Gigg & Gigg *J Chem Soc C* 1865 1967, *Beilstein* **6** IV 2247.]

**2-Benzylphenol** [28994-41-4] **M 184.2, m 54.5°, b 312°/760mm, 175°/18mm,  $\text{pK}_{\text{Est}} \sim 10.0$** . Distil 2-benzylphenol in a vacuum and recrystallise it from EtOH. It has a stable form with **m  $\sim$ 52°** and an unstable form with **m 21°**. [*Beilstein* **6** H 675, **6** IV 4628.]

**4-Benzylphenol ( $\alpha$ -phenyl-*p*-cresol)** [101-53-1] **M 184.2, m 84°,  $\text{pK}_{\text{Est}} \sim 10.2$** . Crystallise 4-benzylphenol

from water. [Beilstein 6 H 675, 6 IV 4628.]

**4-N-Benzylsulfanilamide (Septazen)** [1709-54-2] M 262.3, m 175°, 178°. Crystallise Septazen from dioxane/H<sub>2</sub>O, EtOH/H<sub>2</sub>O or Me<sub>2</sub>CO (m 174.5-175.8°). Its solubility in H<sub>2</sub>O at 37° is 0.03-0.43mg/100ml. [Beilstein 14 III 2026.]

**Benzylthiocyanate** [3012-37-1] M 149.2, m 43°, b 256°(dec). Crystallise the thiocyanate from EtOH or aqueous EtOH. [Beilstein 6 H 460, 6 IV 2680.]

**Benzyl toluene-*p*-sulfonate** [1024-41-5] M 162.3, m 58°, 58.5-58.8°. Crystallise the ester from petroleum ether (b 40-60°), CHCl<sub>3</sub>/hexane or Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub>. Dry it *in vacuo* but NOT in a desiccator over CaCl<sub>2</sub> as it causes hydrolysis of the ester. [Emmons & Ferris *J Am Chem Soc* 75 2257 1953, Beilstein 11 II 48, 11 III 207, 11 IV 273.]

**Benzyltributylammonium bromide** [25316-59-0] M 356.4, m 169-171°, 174-175°. Recrystallise the bromide from EtOAc/EtOH and EtOH/Et<sub>2</sub>O. [Kantor & Hauser *J Am Chem Soc* 73 4122 1951, Petersen et al. *J Am Chem Soc* 81 3264 1959, Beilstein 12 IV 2166.]

**Benzyl 2,2,2-trichloroacetimidate** [81927-55-1] M 252.5, b 106°/0.5mm, m 3°, d<sub>4</sub><sup>20</sup> 1.349, n<sub>D</sub><sup>20</sup> 1.545. Purify the imidate by distillation to remove up to 1% of PhCH<sub>2</sub>OH as stabiliser. A solution in hexane can be stored for up to 2 months without decomposition. It is *hygroscopic* and has to be stored dry. [Wessel et al. *J Chem Soc, Perkin Trans 1* 2247 1985, Beilstein 6 IV 2265.]

**Benzyltrimethylammonium chloride** [56-93-9] M 185.7, m 238-239°(dec). A 60% aqueous solution of the salt is evaporated to dryness under a vacuum on a steam bath, and then left in a vacuum desiccator containing a suitable drying agent. The solid residue is dissolved in a small volume of boiling absolute EtOH and precipitated by adding an equal volume of diethyl ether with cooling. After washing, the precipitate is dried under a vacuum [Karusch *J Am Chem Soc* 73 1246 1951]. [Beilstein 12 IV 2162.]

**Benzyltrimethylammonium hydroxide (Triton B)** [100-85-6] M 167.3, d 0.91. A 38% solution (as supplied) is decolorised (charcoal), then evaporated under reduced pressure to a syrup, with final drying at 75°/1mm pressure. The *anhydrous* base is obtained by prolonged drying over P<sub>2</sub>O<sub>5</sub> in a vacuum desiccator. [Beilstein 12 IV 2162.]

**Bibenzyl (1,2-diphenylethane, dibenzyl)** [103-29-7] M 182.3, m 50-53°, 52.0-52.5°, 52.5-53.5°. Crystallise bibenzyl from hexane, MeOH, or 95% EtOH. It has also been sublimed under vacuum, and further purified by percolation through columns of silica gel and activated alumina. It is prepared by reduction of benzoin or benzyl [Clemmensen *Chem Ber* 47 688 1914], but is best obtained by catalytic reduction of stilbene. Thus a mixture of stilbene (5g, see [103-30-0]) in dioxane (200ml) and Raney Ni (22g) is refluxed with string for 24 hours. The catalyst is filtered off, the filtrate is evaporated *in vacuo*, and the residue is recrystallised from MeOH to give an 80% yield of bibenzyl m 52.0-52.5°. Its FT-IR (melt) has  $\nu_{\max}$  at 3062.6, 1601.7, 1494.9, 1030.2, 906.0, 752.1, 696.9, 580.2 and 519.3 cm<sup>-1</sup>; its <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 7.27 (t, 4H, arom-H), 7.18 (t, 6H, arom-H) and 2.90 (s, 4H, benzylic-H); and its <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, CDCl<sub>3</sub> as internal standard with  $\delta$  at ~77.0) has  $\delta$  at 141.7, 128.36, 128.24, 125.83 and 37.89. [Kleiderer & Kornfeld *J Org Chem* 13 455 1948.] [Beilstein 5 IV 1868.]

(±)-**1,1'-Bi-(2-naphthol) [1,1'-di-(2-naphthol), BINOL]** [602-09-5, 41024-90-2] M 286.3, m 215-217°, 218°, pK<sub>Est(1)</sub> ~7.1, pK<sub>Est(2)</sub> ~11.2. Crystallise the binaphthol from toluene or \*benzene (10ml/g). When crystallised from chlorobenzene it has m 238°. Its solubility in dioxane is 5%. [Beilstein 6 IV 7020.]

**1,1'-Bi-(2-naphthol) [1,1'-di-(2-naphthol)]** [*R*-(+)- 18531-94-7], [*S*-(-)- 18531-99-2] M 286.3, m 207.5-208.5°, 209-211°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> (+) and (-) 37.4.0° (c 0.5, THF), [ $\alpha$ ]<sub>D</sub><sup>25</sup> (+) and (-) 51° (c 0.1, THF), pK as above. Dissolve it in cold 2.5N NaOH, extract with CH<sub>2</sub>Cl<sub>2</sub>, and acidify with 5% HCl. Collect the white precipitate and

recrystallise it from aqueous EtOH and dry it in a vacuum [Akimoto & Yamada *Tetrahedron* **27** 5999 1971]. It is optically stable in dioxane-water (100°/24 hours). *Racemisation*: 72% in 1.2N HCl at 100°/24 hours and 68% in 0.67M KOH in BuOH at 118°/23 hours [Kyba et al. *J Am Chem Soc* **95** 2693 1973]. It has also been crystallised from \*C<sub>6</sub>H<sub>6</sub> (solubility is 1%) using Norite or aqueous EtOH after chromatography through silica gel, eluting with Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>. [Kyba et al. *J Org Chem* **42** 4173 1977; see also Brussee & Jansen *Tetrahedron Lett* **24** 3261 1983, Akimoto & Yamada *Tetrahedron* **27** 5999 1971, *Beilstein* **6** IV 7020.]

**1,1'-Binaphthyl** [(±)- 32507-32-7 and 604-53-5, R(-)- 24161-30-6, S(+)- 734-77-0] **M 254.3, m 145°, 159°, b ~240°/13mm, (±, 2 forms), 153-154°, 154°, (+ and -), [α]<sub>D</sub><sup>20</sup> (-) and (+) ~220° (\*C<sub>6</sub>H<sub>6</sub>).** Purify 1,1'-binaphthyl through a silica gel column with Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub> [or Al<sub>2</sub>O<sub>3</sub> with 10% \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 30-60°)] and recrystallise it from EtOH, pentane, or slow evaporation of \*C<sub>6</sub>H<sub>6</sub>, Me<sub>2</sub>CO or Et<sub>2</sub>O solutions. Half life ~10 hours at 25° in various solvents. [Wilson & Pincock *J Am Chem Soc* **97** 1474 1975, Akimoto & Yamada *Tetrahedron* **27** 5999 1971, *Beilstein* **5** I 358, **5** II 642, **5** III 2465, **5** IV 2634.]

**2,2'-Binaphthyl (β, β'-binaphthyl)** [61-78-2] **M 254.3, m 188°.** Crystallise the 2,2'-binaphthyl from \*C<sub>6</sub>H<sub>6</sub>, or Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> (**m 187-189°**). The 2,4,7-trinitrofluorenone complex forms orange-red needles from EtOH/\*C<sub>6</sub>H<sub>6</sub> (**m 170.6-171°**). [*Beilstein* **5** H 727, **5** I 359, **5** II 643, **5** III 2467, **5** IV 2636.]

**Biphenyl** [92-52-4] **M 154.2, m 68-75°, 70-71°, b 112°/7mm, 255°/760mm, d<sub>4</sub><sup>20</sup> 0.992.** Crystallise biphenyl from EtOH, MeOH, aqueous MeOH, petroleum ether (b 40-60°) or glacial acetic acid. Free it from polar impurities by passage through an alumina column in \*benzene, followed by evaporation. The residue has been purified by distillation in a vacuum and by zone refining. Treatment with maleic anhydride removes anthracene-like impurities. It has been recrystallised from EtOH followed by repeated vacuum sublimation and passage through a zone refiner. [Taliani & Bree *J Phys Chem* **88** 2351 1984, *Beilstein* **5** H 576, **5** I 271, **5** II 479, **5** III 1726, **5** IV 1807.]

**4-Biphenylcarbonyl chloride** [14002-51-8] **M 216.7, m 114-115°.** Dissolve the carbonyl chloride in a large volume of petroleum ether (10 x, b 50-70°), filter it through a short column of neutral alumina, evaporate to dryness *in vacuo* and recrystallise it from petroleum ether (b 60-80°). [*Beilstein* **9** IV 2480.]  
**LACHRYMATORY.**

**Biphenyl-2-carboxylic (2-phenylbenzoic) acid** [947-84-2] **M 198.2, m 114°, b 343-344°, pK<sup>25</sup> 3.46.** Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or aqueous EtOH. [*Beilstein* **9** IV 2472.]

**Biphenyl-4-carboxylic (4-phenylbenzoic) acid** [92-92-2] **M 198.2, m 228°, pK<sup>25</sup> 5.66 (in 50% 2-butoxyethanol).** Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or aqueous EtOH. [*Beilstein* **9** IV 2479.]

**2,4'-Biphenyldiamine** [492-17-1] **M 184.2, m 45°, b 363°/760mm, pK<sub>Est(1)</sub> ~4.8, pK<sub>Est(2)</sub> ~3.9.** Crystallise the diamine from aqueous EtOH or petroleum ether (**m 54-54.5°**). [*Beilstein* **9** III 416, **9** IV 360.]

**Biphenyl-4,4'-dicarboxylic acid** [787-70-2] **M 242.2, m >300°, pK<sub>Est(1)</sub> ~3.5, pK<sub>Est(2)</sub> ~4.3.** The dicarboxylic acid is a white amorphous or microcrystalline substance which does not melt or sublime. It is best purified by precipitation of an aqueous alkaline solution with mineral acid, washing well with H<sub>2</sub>O and drying *in vacuo* at 100°. It is characterised by conversion to *diphenyl-4,4'-dicarbonyl chloride* (with PCI<sub>5</sub>) [Work *J Chem Soc* 1317 1940], or by phase transfer catalysis with SOCl<sub>2</sub> + BuEt<sub>3</sub>N<sup>+</sup>Cl<sup>-</sup> in 1,2-dichloroethane [Burdett *Synthesis* 441 1991] which crystallises from \*C<sub>6</sub>H<sub>6</sub> with **m 184°**. The di-acid chloride gives the *dimethyl ester* with MeOH, and has **m 215-217°** (plates from MeOH, **m**'s of 214° and 224° were also reported). The *diethyl ester* is similarly prepared with EtOH and has **m 122°** (from EtOH). [*Beilstein* **9** II 665, **9** III 4519, **9** IV 3563.]

**Biphenylene** [259-79-0] **M 152.2, m 111°.** Biphenylene forms yellow crystals from cyclohexane, MeOH (**m 110-111°**) or EtOH (**m 111-112°**). It sublimes *in vacuo*. The 2,4,7-trinitrofluorenone complex has **m 154°** and the *picrate* gives red needles **m 122°** from EtOH. [*Beilstein* **5** I 298, **5** II 530, **5** III 1935, **5** IV 2137.]

**(±)-α-(4-Biphenyl)butyric acid** [959-10-4] **M 240.3, m 124-125°, pK<sub>Est</sub> ~4.5.** Crystallise the acid from

MeOH, petroleum ether or AcOH (**m** 123-125°). [*Beilstein* 9 III 3370, 9 IV 2558.]

$\gamma$ -(4-Biphenyl)butyric acid [6057-60-9] **M** 240.3, **m** 118°, 120-121°, **pK<sub>Est</sub>** ~4.8. Crystallise the acid from MeOH (**m** 118°) or \*C<sub>6</sub>H<sub>6</sub> (**m** 118-119°). [*Beilstein* 9 I 290, 9 III 3370, 9 IV 2558.]

Bis-(*p*-bromophenyl) ether [53563-56-7] **M** 328.0, **m** 60.1-61.7°. Crystallise the ether twice from EtOH, petroleum ether, once from \*benzene and dry it *in vacuo* [Purcell & Smith *J Am Chem Soc* 83 1063 1961]. [*Beilstein* 6 III 745, 9 IV 1048.]

2*R*,3*R*-(+)-1,4-Bis-(4-chlorobenzoyloxy)-2,3-butanediol [85362-86-3] and 2*S*,3*S*-(-)-1,4-Bis-(4-chlorobenzoyloxy)-2,3-butanediol [85362-85-2] **M** 371.3, **m** 76-77°,  $[\alpha]_D^{20}$  (+) and (-) 6.4° (c 3.11 CHCl<sub>3</sub>). Recrystallise the diol from toluene-hexane. [Tamoto & Sugimori *Tetrahedron Lett* 23 4107 1982, Tamoto *Tetrahedron* 40 4617 1984.]

*N,N*-Bis-(2-chloroethyl) 2-naphthylamine (chlornaphthazine) [494-03-1] **M** 268.3, **m** 54-56°, **b** 210°/5mm, **pK<sub>Est</sub>** ~5.3. Crystallise it from petroleum ether. At 15° it is soluble in EtOH (3.2%), Et<sub>2</sub>O (50%), Me<sub>2</sub>CO (84%) and \*C<sub>6</sub>H<sub>6</sub> (80%). [*Beilstein* 12 III 2996, 12 IV 3126.] **CARCINOGENIC**.

1,4-Bis-(chloromethyl)durene (1,4-bischloromethyl-2,3,5,6-tetramethylbenzene) [3022-16-0] **M** 231.2, **m** 197-198°. Crystallise it three times from \*C<sub>6</sub>H<sub>6</sub> (**m** 193-194°) or petroleum ether (**m** 195-196°), then dry it *in vacuo* in a drying pistol. [Fuson et al. *J Am Chem Soc* 75 5952 1953, *Beilstein* 5 IV 1140.]

2,2-Bis-(*p*-chlorophenyl)-1,1-dichloroethane (*p,p'*-DDD) [72-54-8] **M** 320.1, **m** 109-111°, 111-112°. Crystallise DDD from EtOH and dry it *in vacuo*. The purity is checked by TLC. [*Beilstein* 5 III 1830.] **TOXIC INSECTICIDE**.

2,2-Bis-(*p*-chlorophenyl)-1,1-dichloroethylene (*p,p'*-DDE) [72-55-9] **M** 318.0, **m** 89-91°. Crystallise DDE from MeOH or EtOH and dry it *in vacuo*. The purity is checked by TLC. [Gätzi & Stambach *Helv Chim Acta* 28 569 1946, *Beilstein* 5 H 639, 5 III 1891.] **POSSIBLE CARCINOGEN**.

2,2-Bis-(4-chlorophenyl)-1,1,1-trichloroethane (*p,p'*-DDT, 1,1,1-trichloro-2,2-bis(*p*-chlorophenyl)ethane) [50-29-3] **M** 354.5, **m** 108.5-109°, 108°. Crystallise DDT from *n*-propyl alcohol (5ml/g), then dry it in air or an air oven at 50-60°. Alternatively, crystallise it from 95% EtOH, and the purity is checked by TLC. [*Beilstein* 5 III 1833.] **TOXIC INSECTICIDE**.

4,4'-Bis-(diethylamino)benzophenone [90-93-7] **M** 324.5, **m** 93-95°, 95-96°, **b** ~300°/10mm, **pK<sub>Est(1)</sub>** ~1.8, **pK<sub>Est(2)</sub>** ~3.3. Crystallise the phenone from EtOH (25ml/g) and dry it under vacuum. Its *picrate* forms yellow needles from EtOH with **m** 178.5°. [*Beilstein* 14 II 59.]

Bis-(4-dimethylaminobenzylidene)benzidine [6001-51-0] **M** 454.5, **m** 318°, **pK<sub>Est</sub>** ~0. Crystallise the benzidine from nitrobenzene. [*Beilstein* 14 H 35.]

Bis-(4-fluoro-3-nitrophenyl) sulfone [312-30-1] **M** 344.3, **m** 193-194°. Recrystallise the sulfone from Me<sub>2</sub>CO and H<sub>2</sub>O (5:1). It should give a yellow colour in aqueous base. [Zahn & Zuber *Chem Ber* 86 172 1953, *Beilstein* 6 IV 172.]

3,4-Bis-(4-hydroxyphenyl)hexane (Hexesterol) [5635-50-7 (no configuration), 84-16-2 (*meso*-3*RS*,4-*SR*)] **M** 270.4, **m** 185-186°, 187°. Free it from diethylstilboestrol by zone refining. Crystallise *meso*-Hexesterol from \*benzene or aqueous EtOH (**m** 185-188°). The *meso*-dibenzoyl derivative has **m** 236-237°. The 3*RS*,4*RS*(±)-*racemate* [5776-72-7] crystallises from petroleum ether, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, Et<sub>2</sub>O/petroleum ether, or MeOH/H<sub>2</sub>O and has **m** 128-129°. The (±)-dibenzoyl derivative has **m** 123-124°. The 3*R*,4*R*(+)-*isomer* [26614-21-1] and 3*S*,4*S*(-)-*isomer* [26614-22-2] crystallise from Et<sub>2</sub>O/petroleum ether with **m** 80-80.5° and have  $[\alpha]_D^{17}$  (+) and (-) 17.7° (c 5, EtOH). Their *dibenzoyl* derivatives have **m** 116.5°. [*Beilstein* 6 III 5503, 6 IV 6761.] They have estrogenic activity where optically active forms are more potent, and they have antineoplastic activity.

[Aboul-Enein et al. *Anal Profiles Drug Subst* **11** 347 1982, *J Am Chem Soc* **65** 4911941.]

**4,4-Bis(4-hydroxyphenyl)valeric acid [diphenolic acid]** [126-00-1] **M 286.3, m 168-171°, 171-172°, pK<sub>Est(1)</sub>~ 4.8 (CO<sub>2</sub>H), pK<sub>Est(2)</sub>~ 7.55 (OH), pK<sub>Est(3)</sub>~9.0 (OH)**. When recrystallised from \*C<sub>6</sub>H<sub>6</sub>, the crystals have 0.5 mol of \*C<sub>6</sub>H<sub>6</sub> (**m** 120-122°), and when recrystallised from toluene, the crystals have 0.5 mol of toluene. Purify the acid by recrystallisation from hot H<sub>2</sub>O. It is soluble in Me<sub>2</sub>CO, AcOH, EtOH, propan-2-ol, methyl ethyl ketone. It can also be recrystallised from AcOH, heptane/Et<sub>2</sub>O or Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>. It has λ<sub>max</sub> at 225 and 279nm in EtOH. The *methyl ester* has **m** 87-89° (aqueous MeOH to give the *trihydrate*). [Bader & Kantowicz *J Am Chem Soc* **76** 4465 1954, *Beilstein* **10** IV 1890.]

**1,4-Bismethylaminoanthraquinone (Disperse Blue 14)** [2475-44-7] **M 266.3, m 220-222°, C.I. 61500, λ<sub>max</sub> 640 (594)nm**. Purify the anthraquinone by thin-layer chromatography on silica gel plates, using toluene/acetone (3:1) as eluent. The main band is scraped off and extracted with MeOH. The solvent is evaporated and the dye is dried in a drying pistol [Land et al. *J Chem Soc, Faraday Trans 1* **72** 2091 1976]. It crystallises from *n*-butanol with **m** 221-222° and has λ<sub>max</sub> at 539 and 644nm (EtOH). [*Beilstein* **14** H 198, **14** III 440, **14** IV 459.]

**Bis-(1-naphthylmethyl)amine** [5798-49-2] **M 329.4, m 62°, 63-64°, pK<sub>Est</sub> ~8.4**. Crystallise the amine from petroleum ether, Et<sub>2</sub>O (**m** 73-74°) or \*C<sub>6</sub>H<sub>6</sub> (**m** 62°). The *hydrochloride* crystallises from H<sub>2</sub>O as needles **m** 239°, and the picrate has **m** 206°(202°). [*Beilstein* **12** II 741, **12** IV 3195.]

**Bis-(4-nitrophenyl) carbonate** [5070-13-3] **M 304.3, m 142-143°**. Dissolve the carbonate in CHCl<sub>3</sub>, wash it with 2N NaOH (3 x) and once with conc HCl, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and crystallise the residue from toluene (authors say prisms from 15 volumes of \*benzene). [Glatthard & Matter *Helv Chim Acta* **46** 795 1963, *Beilstein* **6** III 820.]

**Bis-(2-nitrophenyl) disulfide** [1155-00-6] **M 308.3, m 192-195°, 195°, 194-197°, 198-199°**. Purify the disulfide by recrystallisation from glacial AcOH or from \*C<sub>6</sub>H<sub>6</sub> and the yellow needles are dried in an oven at 100° until the odour of the solvent is absent. It is sparingly soluble in EtOH and Me<sub>2</sub>CO. [Bogert & Stull *Org Synth Coll Vol I* 220 1941, Bauer & Cymerman *J Chem Soc* 3434 1949, *Beilstein* **6** IV 1672.]

**Bis-(4-nitrophenyl) ether** [101-63-3] **M 260.2, m 142-143°, 144.4-144.7°, 147-148°**. Crystallise the ether twice from \*C<sub>6</sub>H<sub>6</sub> or petroleum ether and dry it *in vacuo*. [*Beilstein* **6** II 822, **6** IV 1290.]

**Bis-(4-nitrophenyl) methane** [1817-74-9] **M 258.2, m 183°, 184°, 187°**. Crystallise the methane twice from \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or AcOH (**m** 188.6-189.6°), and dry it *in vacuo*. [*Beilstein* **5** III 1797, **5** IV 1853.]

**Bis-(trifluoroacetoxy)iodobenzene (BTI)** [2712-78-9] **M 430.0, m 112-114°(dec), 120-121°, 124-126°**. Crystallise the iodo compound from warm trifluoroacetic acid and dry it over NaOH pellets. Recrystallise it also from Me<sub>2</sub>CO/petroleum ether. Its melting point depends on the heating rate. [Spyroudis & Varvoglis *Synthesis* 445 1975, application: Almond et al. *Org Synth* **66** 132 1988.]

***N*-BOC-1,2-phenylenediamine ([2-aminophenyl]carbamic acid, *tert*-butyl ester)** [146651-75-4] **M 208.3, m 109-114°**. Purify the ester by crystallisation from CHCl<sub>3</sub>/hexane (1:1, v/v) and dry it *in vacuo*. [Seto et al. *J Am Chem Soc* **115** 1321 1993, Seto et al. *J Am Chem Soc* **127** 11442 2005.]

**Brilliant Green (4-dimethylaminotriphenyl carbinol)** [633-03-4] **M 482.7, m 209-211°(dec), pK<sup>25</sup> 4.75**: Purify the dye by precipitating the *perchlorate* from aqueous solution (0.3%) after filtering, heating to 75° and adjusting to pH 1-2. Recrystallise it from EtOH/water (1:4) [Kerr & Gregory *Analyst (London)* **94** 1036 1969]. [*Beilstein* **13** IV 2281.]

**4-Bromoacetanilide** [103-88-8] **M 214.1, m 167°**. Crystallise the anilide from aqueous MeOH or EtOH. Purify it by zone refining. [*Beilstein* **12** IV 1504.]

**4-Bromoacetophenone** [99-90-1] M 199.1, m 54°. Crystallise it from EtOH, MeOH or from petroleum ether (b 80-100°). [Tanner *J Org Chem* 52 2142 1987, *Beilstein* 7 IV 647.]

**ω-Bromoacetophenone (phenacyl bromide)** [70-11-1] M 199.1, m 57-58°. Crystallise the bromide from EtOH, MeOH or petroleum ether (b 80-100°). [Tanner *J Org Chem* 52 2142 1987, *Beilstein* 7 IV 649.]

**4-Bromoaniline** [106-40-1] M 172.0, m 66°, pK<sup>25</sup> 3.86. Crystallise the aniline (with appreciable loss) from aqueous EtOH. The *benzoyl* derivative has m 204° (from EtOH). [*Beilstein* 12 IV 1497.]

**2-Bromoanisole** [578-57-4] M 187.0, f 2.5°, b 124°/40mm, d<sub>4</sub><sup>20</sup> 1.513, n<sub>D</sub><sup>25</sup> 1.5717. Crystallise the anisole by repeated partial freezing, then distil it under reduced pressure. [*Beilstein* 6 IV 1037.]

**4-Bromoanisole** [104-92-7] M 187.0, f 13.4°, b 99-100°/18mm, 124°/40mm, d<sub>4</sub><sup>20</sup> 1.495, n<sub>D</sub><sup>25</sup> 1.5617. Crystallise the anisole by repeated partial freezing, then distil it under reduced pressure. [*Beilstein* 6 III 741, 6 IV 1044.]

**9-Bromoanthracene** [1564-64-3] M 257.1, m 98-100°. Crystallise 9-bromoanthracene from MeOH or EtOH followed by sublimation *in vacuo*. [Masnori et al. *J Am Chem Soc* 108 126 1986, *Beilstein* 5 IV 2295.]

**4-Bromobenzal diacetate** [55605-27-1] M 287.1, m 95°. Crystallise the diacetate from hot EtOH (3ml/g). [Liebermann & Connor *Org Synth Coll Vol II* 442 1948, *Beilstein* 7 II 182, 7 IV 579.]

**Bromobenzene** [108-86-1] M 157.0, b 155.9°, d<sub>4</sub><sup>20</sup> 1.495, n<sub>D</sub><sup>20</sup> 1.5588, n<sub>D</sub><sup>15</sup> 1.56252. Wash bromobenzene vigorously with conc H<sub>2</sub>SO<sub>4</sub>, then 10% NaOH or NaHCO<sub>3</sub> solutions, and H<sub>2</sub>O. Dry it with CaCl<sub>2</sub> or Na<sub>2</sub>SO<sub>4</sub>, or pass it through activated alumina, before refluxing with, and distilling from, CaH<sub>2</sub>, using a glass helix-packed column. [*Beilstein* 5 IV 670.]

**4-Bromobenzene diazonium tetrafluoroborate** [673-40-5] M 270.8, m 133°(dec), 135-140° (dec), 135°(dec). Wash the salt with Et<sub>2</sub>O until the wash is colourless and allow it to dry by blowing N<sub>2</sub> over it. Store it at 0-4° in the dark. [Schiemann & Pillarsky *Chem Ber* 64 1340 1931, *Beilstein* 16 III 517.]

**4-Bromobenzenesulfonyl chloride** [98-58-8] M 255.5, m 73-75°, 74.3-75.1, 75-76°, 77°, b 153°/15mm, 150.6°/13mm. Wash the sulfonyl chloride with cold water, dry and recrystallise it from petroleum ether, or from ethyl ether cooled in powdered Dry-ice after the ether solution had been washed with 10% NaOH until colourless, then dry it with anhydrous Na<sub>2</sub>SO<sub>4</sub>. Alternatively, dissolve it in CHCl<sub>3</sub>, wash it with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and recrystallise it. [Huntress & Carten *J Am Chem Soc* 62 511 1940.] Test for the SO<sub>2</sub>Cl group by dissolving it in EtOH and boiling with NH<sub>4</sub>CNS whereby a yellow amorphous precipitate forms on cooling. [*Beilstein* 11 IV 162.]

**2-Bromobenzonitrile** [2042-37-7] M 182.0, m 52-54°, 55.8°, 56°, 53-57°, b 251-253°/754mm. The nitrile is prepared from 2-bromobenzoic acid by treatment with SOCl<sub>2</sub> (23 hours boiling) and distilling to give 2-bromobenzoyl chloride, b 120-126°/15mm, n<sub>D</sub><sup>20</sup> 1.5925, which on treatment with 12 equivalents of NH<sub>3</sub> gives a 98% yield of 2-bromobenzamide m 159.5-161.5° [4001-73-4]. By boiling this amide with excess of SOCl<sub>2</sub> for 17 hours followed by evaporation and steam distillation gives an 84% yield of the nitrile that forms needles when crystallised from H<sub>2</sub>O. [Lutz et al. *J Org Chem* 12 666 1947.] It has also been prepared by the method of Miller [*Org Synth Coll Vol* 3 646, 648 1955] in which the amide is treated with a large excess of POCl<sub>3</sub> in the presence of sodium metabisulfite (Na<sub>2</sub>S<sub>2</sub>O<sub>5</sub>) to give the nitrile (m 55-55.5°) [Herbst & Wilson *J Org Chem* 22 1142 1957]. [*Beilstein* 9 H 348, 9 II 232, 9 III 1387, 9 IV 1013.]

**2-Bromobenzoic acid (o-bromobenzoic acid)** [88-65-3] M 201.0, m 148.9°, 150°, pK<sup>20</sup> 2.88. Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub> or MeOH. The *anilide* has m 141° (from EtOH/H<sub>2</sub>O). [*Beilstein* 9 IV 1011.]

**3-Bromobenzoic acid (m-bromobenzoic acid)** [585-76-2] M 201.0, m 155°, pK<sup>25</sup> 3.81. Crystallise the acid from acetone/water, MeOH or acetic acid. The *anilide* has m 137° (from EtOH/H<sub>2</sub>O). [*Beilstein* 9 IV 1013.]

**4-Bromobenzoic acid** (*p*-bromobenzoic acid) [586-76-5] **M 201.0, m 251-252°, 254-256°, 257-258°,  $n_D^{25}$  1.396.** Crystallise the acid from MeOH, or MeOH/water mixture, 90% EtOH and Et<sub>2</sub>O. The *methyl ester* has **m 81°** from Et<sub>2</sub>O or dilute MeOH. The *anilide* has **m 197°** (from EtOH). [Male & Thorp *J Am Chem Soc* **35** 269 1913, Lamneck *J Am Chem Soc* **76** 406 1954, Vandenbelt et al. *Anal Chem* **26** 926 1954, *Beilstein* **9** IV 1017.]

***p*-Bromobenzophenone** [90-90-4] **M 261.1, m 81°, 81-82°.** Crystallise the phenone from EtOH. The *2,4-dinitrophenylhydrazone* forms orange-red leaflets from dioxane/EtOH with **m 207-209°**. [Allen & Van Allan *J Am Chem Soc* **66** 7 1944, *Beilstein* **7** H 422, **7** III 2079, **7** IV 1378.]

**4-Bromobenzoyl acetonitrile** [4592-94-3] **M 224.1, m 160-164°, 162.4-163.4°.** The nitrile is purified by dissolving in slightly alkaline H<sub>2</sub>O (charcoal), filtering and acidifying with HCl to give colourless needles (**m 162-163°**). It recrystallises from EtOH. With Me<sub>2</sub>SO<sub>4</sub>/KOH at 130° it gives *4-bromo-β-methoxycinnamyl nitrile* **m 58.5-59.5°** (from high boiling petroleum ether) [Fuson & Wolf *J Am Chem Soc* **61** 1940 1939, Grathaus & Dains *J Am Chem Soc* **58** 1334 1936]. [*Beilstein* **10** III 2998.]

***p*-Bromobenzoyl chloride** [586-75-4] **M 219.5, m 36-39°, 39.8°, 41°, b 62°/0.1mm, 104.5°/6mm, 126.4-127.2°/14mm.** Check IR of a film to see if OH bands are present. If absent then recrystallise from petroleum ether and dry it *in vacuo*. If OH bands are weak, then distil it *in vacuo* and recrystallise if necessary. If OH bands are very strong, then treat with an equal volume of redistilled SOCl<sub>2</sub> reflux for 2 hours, then evaporate excess of SOCl<sub>2</sub> and distil the residual oil or low melting solid. Store it in the dark away from moisture. **LACHRYMATORY.** [Martin & Partington *J Chem Soc* 1175 1936, *Beilstein* **9** IV 1023.]

***p*-Bromobenzyl bromide** [589-15-1] **M 249.9, m 60-61°.** Crystallise the bromide from EtOH. [*Beilstein* **5** IV 836.] **LACHRYMATORY.**

***p*-Bromobenzyl chloride** [589-17-3] **M 205.5, m 40-41°, b 105-115°/12mm.** Crystallise the chloride from EtOH and distil it in a vacuum. [*Beilstein* **5** IV 832.] **LACHRYMATORY.**

**2-Bromobiphenyl** [2052-07-5] **M 233.1, m 1.5-2.0°, b 140°/11mm, 148-150°/10mm, 297-298°/atm,  $d_4^{25}$  1.352,  $n_D^{20}$  1.6248.** 2-Bromobiphenyl is prepared from 2-aminobiphenyl (12g) in hot constant-boiling aqueous HBr (21g, 3-equivalents) by diazotisation at 5° with a solution of NaNO<sub>2</sub> (6g) in H<sub>2</sub>O (10ml), then excess of HNO<sub>2</sub> is removed at the end of the reaction by addition of urea with stirring for 20 minutes. This solution is added dropwise into a solution of CuBr (from 15g of CuSO<sub>4</sub>) in constant-boiling aqueous HBr (20ml) with stirring, set aside for 20 minutes, after which the brown complex is decomposed by heating on a steam bath for 2 hours. The dark oil is extracted into Et<sub>2</sub>O, the extract is filtered through glass wool, washed with H<sub>2</sub>O, dilute NaOH, H<sub>2</sub>O again, dried (CaCl<sub>2</sub>), filtered, evaporated and the residual oil is distilled to give a pale yellow oil (7g, b 146-152°/12mm). Further purification can be achieved by it dissolving in \*C<sub>6</sub>H<sub>6</sub>, passing through an alumina column and eluting with \*C<sub>6</sub>H<sub>6</sub>, evaporating to dryness and distilling the residual oil (b 148-150°/10mm) to yield pure colourless 2-bromobiphenyl. This has been recrystallised from pentane at -40° to give a liquid with a freezing point of 1.5-2.0°. [de la Mare & Hassan *J Chem Soc* 3004 1957, Augood et al. *J Chem Soc* 3412 1953, NMR: Brownstein *J Am Chem Soc* **80** 2300 1958, *Beilstein* **5** H 580, **5** II 485, **5** III 1742, **5** IV 1818.] It is used for the preparation of John-Phos and related catalytic ligands (see Chapter 7, Catalysis—Part 2).

**3-Bromobiphenyl** [2113-57-7] **M 233.1, b 103-104°/0.2mm, 110°/1mm, 158-167°/11mm, 169-173°/17mm,  $d_4^{25}$  1.3976,  $n_D^{20}$  1.6380.** 3-Bromobiphenyl can be prepared by bromination of 2-acetamidobiphenyl (m 118-119° from aqueous AcOH) with one molecular equivalent of Br<sub>2</sub> in AcOH; the monobromo derivative (m 127-127.5° from EtOH) is hydrolysed with 95% EtOH/conc HCl (1.4:1, 4 hours reflux) poured into excess of H<sub>2</sub>O, basified and filtered off to give 2-amino-5-bromobiphenyl (m 53-56°). The base is subsequently diazotised and the diazonium salt is deaminated by warming with Cu bronze or with hypophosphorous acid [cf Kornblum in *Organic Reactions* J. Wiley & Sons NY, **II** 294 1944] to yield crude 3-bromobiphenyl which is isolated by steam distillation, extraction of the distillate with \*C<sub>6</sub>H<sub>6</sub>, washing the extract successively with, dilute NaOH,

H<sub>2</sub>O, conc H<sub>2</sub>SO<sub>4</sub>, H<sub>2</sub>O, drying (K<sub>2</sub>CO<sub>3</sub>), evaporating and distilling. This can be purified further by passing a solution of it in \*C<sub>6</sub>H<sub>6</sub> through an Al<sub>2</sub>O<sub>3</sub> column, eluting with \*C<sub>6</sub>H<sub>6</sub>, and the combined eluates are distilled. The distillate is dissolved in petroleum ether (b 40-60°), washed several times with concentrated H<sub>2</sub>SO<sub>4</sub>, H<sub>2</sub>O again, aqueous NaHCO<sub>3</sub>, H<sub>2</sub>O again, dried (CaCl<sub>2</sub>), filtered, the solvent is evaporated off and the residue is fractionally distilled to give pure *3-bromobiphenyl*. [de la Mare & Hassan *J Chem Soc* 3004 1957, Lichtin & Leftin *J Am Chem Soc* 74 4207 1952, Huber et al. *J Am Chem Soc* 46 1111 1946]. It has also been prepared in low yield (13%) from diazotized *m*-bromoaniline and \*C<sub>6</sub>H<sub>6</sub> according to Gomberg and purified as above [Marvel et al. *J Am Chem Soc* 61 77 1939, also see 4-bromobiphenyl below]. [Beilstein 5 II 485, 5 III 1742, 5 IV 1818.]

**4-Bromobiphenyl** [92-66-0] M 233.1, m 82-86°, 88.8-89.2°, 89.5°, 90°, b 170-175°/8mm, 310°/atm. This biphenyl is prepared by the method of Gomberg & Bachmann [*Org Synth* 1 113 1944] from diazotised 4-bromoaniline and benzene, then extract into Et<sub>2</sub>O, wash with base, acid, brine, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and recrystallise to constant melting point from absolute EtOH (89.7°, after drying *in vacuo*). [Hey et al. *J Chem Soc* 1284 1940; Augood et al. *J Chem Soc* 3412 1953, Beilstein 5 H 580, 5 I 275, 5 II 485, 5 III 1742, 5 IV 1819.]

**1-Bromo-2-chlorobenzene** [694-80-4] M 191.5, m -12°, b 79-82°/14mm, 200-202°/atm, 204°/atm, d<sub>4</sub><sup>25</sup> 1.638, n<sub>D</sub><sup>20</sup> 1.582. *o*-Chlorobromobenzene has been prepared in at least three different ways. It is obtained in 54% yield by carrying out a Sandmeyer reaction from *o*-chloroaniline *via* diazotisation (using NaNO<sub>2</sub>/HBr) and decomposing it with CuBr in 48% aqueous HBr. Steam distil off the product, basify the distillate with 20% aqueous NaOH (extract it into Et<sub>2</sub>O if quantities are small), dry the organic layer (Na<sub>2</sub>SO<sub>4</sub>), filter and distil [Fry & Grote *J Am Chem Soc* 48 710, 1926]. The second procedure involves adding Br (32g) in CCl<sub>4</sub> (50ml) during 20 minutes to refluxing silver 2-chlorobenzoate in CCl<sub>4</sub> (250ml), boil for a further 30 minutes to complete the decarboxylation, filter off the AgBr, wash the filtrate with aqueous Na<sub>2</sub>SO<sub>3</sub> (until red colour of Br is discharged), aqueous 0.5M Na<sub>2</sub>CO<sub>3</sub> (2 x 25ml), evaporate the solvent and distil bromochlorobenzene (38-48% yield) [Dauben & Tilles *J Am Chem Soc* 72 3185 1950, Barnes & Prochaska *J Am Chem Soc* 72 3188 1950]. The third procedure involves a phase transfer catalytic radical bromination. *o*-Chlorobenzene free radical is generated from the corresponding *o*-benzenediazonium tetrafluoroborate with KOAc in tetrahydrofuran and a catalytic amount of 18-crown-6 using CBrCl<sub>3</sub> as brominating agent to produce *o*-bromochlorobenzene in 70% yield, a great improvement on the Sandmeyer method [Korzeniowski & Gokel *Tetrahedron Lett* 3519 1977]. [Beilstein 5 H 209, 5 I 115, 5 II 161, 5 III 562, 5 IV 680.]

**4-Bromo-4'-chlorobenzophenone** [27428-57-5] M 295.6, m 150°. Crystallise the phenone from EtOH or \*C<sub>6</sub>H<sub>6</sub> and further purify it by zone refining (100 passes) [Grove & Turner *J Chem Soc* 509 1929, Lin & Hanson *J Phys Chem* 91 2279 1987]. [Beilstein 7 II 360, 7 III 2081.]

**Bromocresol Green (3',3'',5',5''-tetrabromo-*m*-cresolsulfonephthalein)** [76-60-8] M 698.0, m 218-219°(dec), 225°(dec), pK<sup>25</sup> 4.51. Crystallise the dye from glacial acetic acid or dissolve it in aqueous 5% NaHCO<sub>3</sub> solution and precipitate it from the hot solution by dropwise addition of aqueous HCl. Repeat this until the UV/VIS-extinction did not increase at λ<sub>max</sub> 423nm. It is an indicator: at pH 3.81 (yellow) and pH 5.4 (blue-green). [Beilstein 19/3 V 460.]

**Bromocresol Purple (5',5''-dibromo-*o*-cresolsulfonephthalein)** [115-40-2] M 540.2, m 241-242°(dec), pK<sub>1</sub> -2.15, pK<sub>2</sub> 6.3. Dissolve the dye in aqueous 5% NaHCO<sub>3</sub> solution and precipitate it from a hot solution by dropwise addition of aqueous HCl. Repeat this until the UV/VIS-extinction did not increase at λ<sub>max</sub> 419nm. It can also be recrystallised from \*benzene. It is an indicator: at pH 5.2 (yellow) and pH 6.8 (purple). [Beilstein 19/3 V 460.]

**2'-Bromo-2,6-dimethoxybiphenyl** [755017-61-9] M 293.2, m 141-142°, 143-146°. This intermediate, which is used for the preparation of the ligand S-Phos, is synthesised by adding *n*-BuLi (9.60ml, 1.6M solution in hexanes, 15.4mmol, 1.2 equivalents) *via* a syringe over 5 minutes to a cold (0°) solution of 1,3-dimethoxybenzene (2.00ml, 15.3mmol, 1.2 equivalents, [151-10-0]) in dry THF (30ml), allowing the tempera-

ture to rise to  $\sim 25^\circ$ , then it is stirred for 5 hours. The mixture is re-cooled ( $0^\circ$ ) and 2-bromochlorobenzene (1.50ml, 12.8mmol, 1.0 equivalents, [694-80-4]) is added dropwise *via* a syringe over 15 minutes while stirring vigorously; and the burgundy coloured solution is stirred for a further 15 minutes at  $0^\circ$ . MeOH (0.25ml) is added *via* syringe to decompose excess BuLi. the whole is evaporated to dryness *in vacuo*, the residue is then stirred with Et<sub>2</sub>O (50ml) and H<sub>2</sub>O (50ml), the layers are separated, the aqueous phase is extracted with Et<sub>2</sub>O (2 x 25ml), the combined Et<sub>2</sub>O solutions are dried (MgSO<sub>4</sub>), filtered and evaporated *in vacuo*. The yellow residue is recrystallised from MeOH to provide the analytically pure biphenyl (3.03g, 81%) as pale yellow crystals. The IR (film) has  $\nu_{\max}$  at 2964, 1584, 1472, 1432, 1248, 1108, 1025, 783 cm<sup>-1</sup>; the <sup>1</sup>H NMR [100MHz, CDCl<sub>3</sub>] has  $\delta$  at 7.69 (dd,  $J = 6.9, 1.1$ Hz, 1H), 7.34-7.40 (m, 2H), 7.20-7.28 (m, 2H), 6.68 (d,  $J = 8.5$ Hz, 2H), 3.76 (s, 6H) from TMS; the <sup>13</sup>C NMR [75MHz, <sup>1</sup>H decoupled, CDCl<sub>3</sub>] has  $\delta$  at 157.8, 136.25, 132.52, 132.47, 129.6, 128.8, 127.1, 125.4, 119.0, 104.2, 56.2 from TMS. [Barder et al. *J Am Chem Soc* **127** 4685 2005.]

***p*-Bromo-*N,N*-dimethylaniline** [586-77-6] **M 200.1, m 55°, b 264°, pK 19.2 (acidic), pK<sup>25</sup> 4.23**. Reflux the aniline for 3 hours with two equivalents of acetic anhydride, then fractionally distil it under reduced pressure. [Beilstein **12** IV 1499.]

**1-Bromo-2,4-dinitrobenzene** [584-48-5] **M 247.0, m 72.5-73°, 75°**. Crystallise it from ethyl ether, isopropyl ether, 80% EtOH or absolute EtOH. [Beilstein **5** III 640, **5** IV 749.]

***N*-(2-Bromoethyl)phthalimide** [574-98-1] **M 254.1, m 81-83°, 82.5-83.5°**. The following is to be carried out in an efficient **FUME HOOD**. Dissolve the compound (180g) in CS<sub>2</sub> (500 ml) by refluxing for 15 minutes (to cause the separation of the most likely impurity, 1,2-diphthalimidoethane), filter and evaporate under reduced pressure. The product forms light tan crystals (**m** 78-80°). Recrystallise it from EtOH (charcoal) [the compound (50g) is dissolved in hot 75% EtOH (200ml), boiled for *ca* 10 minutes, carbon is added (5g, Norite), filter and cool to  $0^\circ$ ], to give white crystals (40g) which can be recrystallised (**m** 80-81°); and further recrystallisation gives **m** 82-83°. [Salzberg & Supniewski *Org Synth Coll Vol I* 119 1932, Landini & Rolla *Synthesis* 389 1976, Beilstein **21/10** V 275.]

**3-Bromo-5-hydroxybenzoic acid** [140472-69-1] **M 217.0, m 233.5°, 237-241°, pK<sub>Est(1)</sub>  $\sim$ 2.3, pK<sub>Est(2)</sub>  $\sim$ 13.0**. The acid crystallises from H<sub>2</sub>O (**m** 238-239°), and with Me<sub>2</sub>SO<sub>4</sub> it yields the *3-methoxy* derivative with **m** 190-191° (from EtOH). [Baddar et al. *J Chem Soc* 469 1955, Beilstein **10** IV 333.]

**2-Bromomethylantraquinone** [7598-10-9] **M 301.1, m 200-202°**. Recrystallise the quinone from AcOH, wash the crystals with a little Et<sub>2</sub>O, dry it in air and then in a vacuum at  $100^\circ$ . It is prepared by bromination of 2-methylantraquinone with Br<sub>2</sub>/PhNO<sub>2</sub> at 145-150°, or *N*-bromosuccinimide in CCl<sub>4</sub> containing a trace of (PhCOO)<sub>2</sub>. [Beilstein **7** IV 2576.]

**2-(Bromomethyl)benzotrile** [22115-41-9] **M 195.1, m 72-73°, 79°, b 152-155°/15mm**. Purify the nitrile by steam distillation. Extract the distillate with Et<sub>2</sub>O, dry the extract (Na<sub>2</sub>SO<sub>4</sub>), evaporate and distil the residue. The solidified distillate can be recrystallised from petroleum ether or cyclohexane. <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 7.8-7.2 (m 4H), 4.62 (s, 2H); the IR has  $\nu_{\text{C}\equiv\text{N}}$  2238 cm<sup>-1</sup>. [Drory *Chem Ber* **24** 2570 1891, Borsche et al. *Chem Ber* **74** 685 1934, Buckley et al. *Aust J Chem* **22** 594 1969, Beilstein **9** III 2312.] **LACHRYMATORY**.

**4-Bromo- $\alpha$ -methylbenzyl alcohol** [( $\pm$ ) 5391-88-8, 25675-29-0, *R*-(+) 76155-78-7, *S*-(-) 100760-04-1] **M 201.1**. The ( $\pm$ )-racemate is purified by distillation in a vacuum (**b** 90°/1mm, 119-121°/7mm, **d** 1.46) and it solidifies on cooling (**m** 36-37°) [Overberger et al. *Org Synth Coll Vol III* 200 1955]. The ( $\pm$ )-*tosyl* derivative [114200-15-6] has **m** 56-57°. The *R*-(+)-*enantiomer* is also purified by distillation in a vacuum (**b** 110°/3mm,  $d_4^{25}$  1.322,  $n_D^{20}$  1.569) and has  $[\alpha]_D^{20} +39^\circ$  (c 1, CHCl<sub>3</sub>),  $+32.9^\circ$  (c 1.39, MeOH). The *S*-(-)-*enantiomer* is similarly purified. [Stein et al. *Can J Chem* **63** 3442 1985, Crevinka et al. *Col Czech Chem Commun* **51** 401 1986, Beilstein **6** II 447.]

**2-(Bromomethyl)-naphthalene** [939-26-4] **M 221.1, m 52-54°, 56°, 56-57°, b 133-136°/0.8mm, 214°/100mm**. Dissolve the bromo compound in toluene, wash it with saturated aqueous NaHCO<sub>3</sub>, dry (Mg

$\text{SO}_2$ , evaporate, fractionally distill the residue and recrystallise the solidified distillate from EtOH. [Chapman & Williams *J Chem Soc* 5044, 1932; Bergmann & Szmaszkowicz *Bull Soc Chim Fr* 20 566 1953; *Beilstein* 5 IV 1692.]

1-Bromonaphthalene [90-11-9] M 207.1, b 118°/6mm,  $d_D^{20}$  1.489. Purify 1-bromonaphthalene by passage through activated alumina, and three vacuum distillations. [*Beilstein* 5 H 547, 5 IV 1665.]

2-Bromonaphthalene [580-12-2] M 207.1, m 59°. Purify 2-bromonaphthalene by fractional elution from a chromatographic column of activated alumina. Crystallise it from EtOH. [*Beilstein* 5 IV 1667.]

1-Bromo-2-naphthol [573-97-7] M 223.1, m 76-78°, 83°, 84°,  $\text{pK}_{\text{Ext}} \sim 8.0$ . Distil the naphthol at 10mm then recrystallise it from \* $\text{C}_6\text{H}_6$  petroleum ether (b 30-60°) m 80-81°. The benzoyl derivative has m 98.5-99.5° (from MeOH). [Hazlet *J Am Chem Soc* 62 2156 1940; *Beilstein* 6 H 650, 6 II 604, 6 III 2994.]

6-Bromo-2-naphthol [15231-91-1] M 223.1, m 122-126°,  $\text{pK}_{\text{Ext}} \sim 9.1$ . Crystallise the naphthol from EtOH or \* $\text{C}_6\text{H}_6$  petroleum ether (m 128°). The benzoyl derivative has m 122° (from EtOH). [Ruggli & Michels *Helv Chim Acta* 14 779 1931; *Beilstein* 6 H 651, 6 II 605, 6 III 2996.]

*o*-Bromo-4-nitroacetophenone [99-81-9] M 244.1, m 98°. Crystallise it from \* $\text{C}_6\text{H}_6$  petroleum ether. [*Beilstein* 7 IV 661.]

*o*-Bromonitrobenzene (2-bromo-1-nitrobenzene) [577-19-5] M 202.1, m 43°. Crystallise it twice from petroleum ether, using charcoal before the first crystallisation. [*Beilstein* 5 III 618, 5 IV 728.]

*m*-Bromonitrobenzene (3-bromo-1-nitrobenzene) [585-79-5] M 202.1, m 55-56°. Crystallise it twice from petroleum ether, using charcoal before the first crystallisation. [*Beilstein* 5 III 618, 5 IV 729.]

*p*-Bromonitrobenzene (4-bromo-1-nitrobenzene) [586-78-7] M 202.1, m 127°. Crystallise it twice from petroleum ether, using charcoal before the first crystallisation. [*Beilstein* 5 III 619, 5 IV 729.]

*p*-Bromophenacyl bromide [99-73-0] M 277.9, m 110-111°. Crystallise the bromide from EtOH (ca 8ml/g). [*Beilstein* 7 IV 652.]

*o*-Bromophenol [95-56-7] M 173.0, b 194°,  $d_D^{20}$  1.490,  $\text{pK}^{25}$  8.45. Purify the phenol by at least two passes through a chromatographic column and distil it. [*Beilstein* 6 IV 1037.]

*p*-Bromophenol [106-41-2] M 173.0, m 64°,  $\text{pK}^{25}$  9.36. Crystallise the phenol from  $\text{CHCl}_3$ ,  $\text{CCl}_4$ , petroleum ether (b 40-60°), or water and dry it at 70° under vacuum for 2 hours. [*Beilstein* 6 IV 1043.]

Bromophenol Blue (3,3',5,5'-tetrabromophenolsulfonephthalein) [115-39-9] M 670.0, m 270-271°(dec), 273°(dec),  $\lambda_{\text{max}}$  422nm,  $\text{pK}^{25}$  3.62 (acidic). Crystallise the dye from \* $\text{C}_6\text{H}_6$  or  $\text{Me}_2\text{CO}$  AcOH, and dry it in air. It is an indicator: at pH 3.0 it is yellow and it is purple at pH 4.6. [*Beilstein* 19/3 V 458.]

(4-Bromophenoxy)acetic acid [1878-91-7] M 231.1, m 157°, 158°,  $\text{pK}^{25}$  3.13. Crystallise the acid from EtOH or  $\text{H}_2\text{O}$  (m 161.4-161.8°). [Hayes & Branch *J Am Chem Soc* 65 1555 1943; *Beilstein* 6 III 747, 6 IV 1052.]

3-(4-Bromophenoxy)propionic acid [93670-18-9] M 247.1, m 146°,  $\text{pK}_{\text{Ext}} \sim 4.2$ . Crystallise the acid from EtOH, MeOH or \* $\text{C}_6\text{H}_6$ /hexane (m 144-145°). [*Beilstein* 6 III 748, 6 IV 1052.]

4-Bromophenylacetic acid [1878-68-8] M 215.1, m 112-113°, 113-115°, 114°,  $\text{pK}^{25}$  4.19. The acid crystallises from  $\text{H}_2\text{O}$  as needles. The acid chloride has b 238-760mm, m 50°, and the amide has m 174-175°. [Dippy & Williams *J Chem Soc* 161 1934 1251 1948; Schwenk & Pala *J Org Chem* 11 798 1946; *Beilstein* 9 III 2275.]

**4-Bromophenylhydrazine** [589-21-9] M 187.1, m 108-109°, pK<sup>20</sup> -5.6 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sup>25</sup> 5.05. Crystallise the hydrazine from H<sub>2</sub>O. The *hydrochloride* crystallises from EtOH/H<sub>2</sub>O with m 213-214°, and the *tosylate* has m 212° (from EtOH). [Beilstein 15 H 434, 15 I 117, 15 II 160, 15 III 289, 15 IV 282.]

**4-Bromophenyl isocyanate** [2492-02-9] M 189.0, m 41-42°, b 158°/14mm. Crystallise the isocyanate from petroleum ether (b 30-40°). It has a pungent odour. [Beilstein 12 H 647, 12 I 321.]

**4-Bromophenyl isothiocyanate** [1985-12-2] M 214.1, m 56-58°. Recrystallise the isothiocyanate from boiling *n*-hexane. Any insoluble material is most probably the corresponding urea. It is also purified by steam distillation, cool the receiver, add NaCl and extract in Et<sub>2</sub>O, wash the extract with N H<sub>2</sub>SO<sub>4</sub>, dry (MgSO<sub>4</sub>), evaporate and recrystallise the residual solid. [Cymerman-Craig et al. *Org Synth Coll Vol IV* 700 1963, cf Dains et al. *Org Synth Coll Vol I* 447 1941, Beilstein 6 IV 1051, 12 II 354, 12 III 1463p, 12 IV 1519.]

***N*-(3-Bromopropyl)phthalimide** [5460-29-7] M 268.1, m 72-74°, 74°. Place it in a Soxhlet and extract it with Et<sub>2</sub>O, whereby the bis-phthalimido impurity is not extracted. Evaporate the Et<sub>2</sub>O and recrystallise the residue from EtOH, aqueous EtOH or petroleum ether. [Gabriel & Weiner *Chem Ber* 21 2669 1888, Gaudry *Can J Chem* 31 1060 1953, Beilstein 21/10 V 1277.]

**Bromopyrogallol Red (5,5'-dibromopyrogallolsulfonephthalein)** [16574-43-9] M 576.2, m 300°, λ<sub>max</sub> 538nm (ε 54,500 H<sub>2</sub>O pH 5.6-7.5), pK<sub>1</sub> 2.9, pK<sub>2</sub> 4.39, pK<sub>3</sub> 9.15, pK<sub>4</sub> 11.72. Crystallise the dye from 50% EtOH, or aqueous alkaline solution followed by acidification. It is a metal chromic indicator. [Suk *Col Czech Chem Commun* 31 3127 1966, Beilstein 19/10 V 226.]

**5-Bromosalicyl hydroxamic acid** [5798-94-7] M 210.1, m 232°(dec), pK<sub>Est(1)</sub>~ 1.5, pK<sub>Est(2)</sub>~ 7.0, pK<sub>Est(3)</sub>~ 8.7. Crystallise the hydroxamic acid from H<sub>2</sub>O (m 249°) or from EtOH (m 232° dec). It sublimes at m 235°. It complexes with metals. [Beilstein 10 IV 221.]

**4-Bromostyrene** [2039-82-9] M 183.1, b 49.5-50°/2.5mm, 87-88°/12mm, 102-104°/20mm, d<sub>4</sub><sup>20</sup> 1.3984, n<sub>D</sub><sup>20</sup> 1.5925. It polymerises above 75° in the presence of benzoyl peroxide. To purify, if it has not gone to a solid resin, dissolve it in Et<sub>2</sub>O, dry (MgSO<sub>4</sub>) and add *ca* 0.1g of 4-*tert*-butylcatechol (polymerisation inhibitor) per 100g of bromostyrene. Filter, evaporate this under reduced pressure (use as high a vacuum as possible) and distil the residue. Store it in dark bottles in the presence of the inhibitor (at above concentration). [Overberger & Saunders *Org Synth Coll Vol III* 204 1955, Beilstein 5 IV 1349.]

**Bromothymol Blue (3',3''-dibromothymolsulfonephthalein)** [76-59-5] M 624.4, m 201-203°, pK<sub>1</sub> -0.66, pK<sub>2</sub> 6.99. Dissolved the dye in aqueous 5% NaHCO<sub>3</sub> solution and precipitate it from the hot solution by dropwise addition of aqueous HCl. Repeat this until the extinction at λ<sub>max</sub> 420 nm does not increase. It is an indicator: aqueous solutions are yellow at pH 6.0, and blue at pH 7.6. [Beilstein 19/3 V 461.]

***o*-Bromotoluene** [95-46-5] M 171.0, m -27°, b 58-60°/10mm, 74°/19mm, 181.7°/760mm, d<sub>4</sub><sup>20</sup> 1.422, n<sub>D</sub><sup>20</sup> 1.556. Fractionally distil it through an efficient column. It can be separated from its isomers by gas chromatography on a column of "Sil-o-cel" firebrick (30-40mesh, 80 parts) coated with 5% (20 parts) of ICI E301 con rubber with N<sub>2</sub> carrier gas at 170°/atm and 100ml/minute and using a conductivity cell detector. [Cowley et al. *J Chem Soc* 1801 1959, Beilstein 5 H 304, 5 I 153, 5 II 234 5 III 704, 5 IV 825.]

***p*-Bromotoluene** [106-38-7] M 171.0, m 28°, b 184°, d<sub>4</sub><sup>20</sup> 1.390. Crystallise it from EtOH [Taylor & Stewart *J Am Chem Soc* 108 6977 1986]. [Beilstein 5 IV 827.]

**α-Bromo-4-toluic acid** [6232-88-8] M 215.1, m 229-230°, pK<sub>Est</sub> ~3.2. Crystallise the acid from Me<sub>2</sub>CO. [Beilstein 9 IV 1745.]

**α-Bromo-*p*-xylene (*p*-methylbenzyl bromide)** [104-81-4] M 185.1, m 35°, b 218-220°/740mm. Crystallise the bromide from EtOH or pentane. [Beilstein 5 H 385, 5 IV 969.]

**2-tert-Butoxycarbonyloxyimino-2-phenylacetoneitrile (BOC-ON)** [58632-95-4] M 246.3, m 87-89°. Triturate the solid with 90% aqueous MeOH, filter, wash with 90% aqueous MeOH and dry it in a vacuum. Recrystallise it from MeOH (needles or plates), but use warm MeOH and cool to crystallise; *do not boil as it decomposes slowly*. Its IR has  $\nu_{\max}$  at 1785 (C=O)  $\text{cm}^{-1}$  and NMR ( $\text{CDCl}_3$ ) usually shows two *tert*-butyl singlets for *syn* and *anti* isomers. Store it in a brown bottle (fridge). It evolves  $\text{CO}_2$  at room temperature (stoppered bottle can explode!), but can be stored over silica gel which may extend its useful life to more than a year. [Itoh et al. *Org Synth* 59 95 1980.]

**4-Butoxyphenylacetic acid** [4547-57-3] M 208.3, m 86-87°, 88.5°,  $\text{pK}_{\text{Est}} \sim 4.4$ . Recrystallise it from petroleum ether (b 40-60°). [McElvain & Carney *J Am Chem Soc* 68 2592 1946, *Beilstein* 10 IV 545.]

**n-Butyl p-aminobenzoate (Butamben)** [94-25-7] M 193.2, m 57-59°, b 174°/8mm,  $\text{pK}_{\text{Est}} \sim 2.5$ . Crystallise Butamben from EtOH. [*Beilstein* 14 IV 1130.]

**tert-Butylammonium bromide** [60469-70-7] M 154.1, m >250°(dec). Recrystallise the salt several times from absolute EtOH or by dissolving in absolute EtOH and adding  $\text{Et}_2\text{O}$  slowly to crystallise the salt. Dry it thoroughly at 105°. [IR: Chenon & Sandorfy *Can J Chem* 36 1181 1958, *Beilstein* 4 IV 659.]

**2-tert-Butylanthracene** [13719-97-6] M 234.3, m 148-149°. Recrystallise the anthracene from EtOH and finally purify it by TLC. [*Beilstein* 5 IV 2364.]

**n-Butylbenzene** [104-51-8] M 134.2, b 183.3°,  $d_4^{20}$  0.860,  $n_D^{20}$  1.4897,  $n_D^{25}$  1.487. Distil butylbenzene from sodium. Wash it with small portions of conc  $\text{H}_2\text{SO}_4$  until the acid is no longer coloured, then with water and aqueous  $\text{Na}_2\text{CO}_3$ . Dry it ( $\text{MgSO}_4$ ), and distil it twice from Na, collecting the middle fraction [Vogel *J Chem Soc* 607 1948]. [*Beilstein* 5 IV 1033.]

**tert-Butylbenzene** [98-06-6] M 134.2, b 169.1°,  $d_4^{20}$  0.867,  $n_D^{20}$  1.493,  $n_D^{25}$  1.490. Wash it with cold conc  $\text{H}_2\text{SO}_4$  until a fresh portion of acid is no longer coloured, then with 10% aqueous NaOH (care-effervescence), followed by distilled water until neutral. Dry it ( $\text{CaSO}_4$ ), and distil it in a glass helices-packed column, taking the middle fraction. [*Beilstein* 5 IV 1045.]

**4-tert-Butyl benzoyl chloride** [1710-98-1] M 196.7, b 135°/10mm, 149.9-150.5°/14mm, 266-268°(dec),  $d_4^{20}$  1.082,  $n_D^{20}$  1.536. Distil it in a vacuum. If IR shows OH group, then treat it with thionyl chloride or oxalyl chloride at *ca* 50° for 30 minutes, evaporate and fractionate it in a vacuum using a short column. Strongly LACHRYMATORY; use a good fume hood. [Fuson & Turnbull *J Am Chem Soc* 71 2544 1949, Tsuno et al. *Bull Chem Soc Jpn* 32 960 1959, Swain et al. *J Am Chem Soc* 72 5433 1950, *Beilstein* 9 III 2526.]

**4-tert-Butylcatechol** [98-29-3] M 166.22, m 47-48°, 55-56°, 75°, b 265°/atm,  $\text{pK}_{\text{Est}(1)} \sim 9.5$ ,  $\text{pK}_{\text{Est}(2)} \sim 13.0$ . Distil it in a vacuum, then recrystallise it from pentane or petroleum ether (or  $^*\text{C}_6\text{H}_6$ ). [*Beilstein* 6 IV 6014.]

**6-tert-Butyl-1-chloro-2-naphthol** [525-27-9] M 232.7, m 76°, b 185°/15mm,  $\text{pK}_{\text{Est}} \sim 8.0$ . Recrystallise the naphthol from petroleum ether. Its *methyl ether* has m 115° (from EtOH/petroleum ether). [Buu-Hoi et al. *J Org Chem* 15 1064 1950, *Beilstein* 6 IV 4367.]

**2-tert-Butyl-4,5-dimethylphenol** [1445-23-4] M 178.3, b 144-150°/20mm,  $\text{pK}_{\text{Est}} \sim 11.8$ . It is obtained by placing 3,4-dimethylphenol (170g, 1.39mol, m  $\sim 65^\circ$ , see [95-65-8]), concentrated  $\text{H}_2\text{SO}_4$  (1.5ml) and a magnetic stirrer bar in a medium pressure vessel, then purging the closed vessel with isobutylene while heating and stirring vigorously at 70° (which keeps the phenol in a moten state) under an isobutylene pressure of 20 psi. The reaction is complete when the liquid stops expanding. The mixture is cooled,  $\text{Et}_2\text{O}$  is added and the mixture is washed with saturated aqueous  $\text{NaHCO}_3$  (3 x 150ml), the combined organic layers are dried ( $\text{MgSO}_4$ ), filtered, the solvent is evaporated and the residual oil is distilled to give 2-tert-butyl-4,5-dimethylphenol, with and the  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 7.02 (s, 1H, aromatic CH), 6.49 (s, 1H, aromatic CH), 2.19 (s, 3H,  $\text{CH}_3$ ), 2.17 (s, 3H,  $\text{CH}_3$ ), 1.40 (s, 9H,  $\text{C}(\text{CH}_3)_3$ ). Note *tert-butyl 3,4-dimethylphenyl ether* is formed instead when lower

temperatures and very small amounts of acid are used [Stevens *J Am Chem Soc* **20** 1232 1955.] [Alexander et al. *J Am Chem Soc* **120** 4041 1998, Albert *J Am Chem Soc* **76** 1983 1954; for acetyl derivative see Fischer & Teo *Can J Chem* **56** 258 1978.]

**2-tert-Butyl hydroquinone** [1948-33-0] **M 166.2, m 125-127°, 127-128°, 129°, pK<sub>Est(1)</sub> ~10.5, pK<sub>Est(2)</sub> ~11.6.** Recrystallise the hydroquinone from H<sub>2</sub>O or MeOH and dry it in a vacuum at 70°. Store it in a dark container. [Stroh et al. *Angew Chem* **69** 699 1957, *Beilstein* **6** IV 6013.]

**2-tert-Butyl-4-methoxyphenol (2-tert-butyl-4-hydroxyanisole)** [121-00-6] **M 180.3, m 64.1°, pK<sub>Est</sub> ~10.8.** Fractionally distil the phenol *in vacuo*, then pass it as a solution in CHCl<sub>3</sub> through alumina, and evaporate the eluate. Recrystallise the residue from petroleum ether. [*Beilstein* **6** IV 6013.]

**p-tert-Butylnitrobenzene** [3282-56-2] **M 179.2, m 28.4°, b 135°/10mm, 140-142°/15mm, n<sub>D</sub><sup>20</sup> 1.5230.** Recrystallise it three times by partially freezing a mixture of the mono-nitro isomers, then recrystallise it twice from MeOH and dry it *in vacuo* [Brown *J Am Chem Soc* **81** 3232 1959]. [*Beilstein* **5** H 418, **5** I 203, **5** II 321, **5** III 943, **5** IV 1052.]

**tert-Butyl perphthalic acid (monoperoxyphthalic acid 1-tert-butyl ester)** [15042-77-0] **M 238.2, m 104-104.5° (dec), pK<sub>Est</sub> ~6.2.** Crystallise the per acid-ester from Et<sub>2</sub>O or Et<sub>2</sub>O/petroleum ether and dry it over H<sub>2</sub>SO<sub>4</sub>. The ester was prepared from *tert*-butylhydroperoxide and phthalic anhydride [Davies et al. *J Chem Soc* 1545 1953]. Possibly **EXPLOSIVE**. [*Beilstein* **9** IV 3260.]

**p-tert-Butylphenol** [98-54-4] **M 150.2, m 99°, pK<sup>25</sup> 10.39.** Crystallise the phenol to constant melting point from petroleum ether (b 60-80°). It sublimes *in vacuo*. Also purify it *via* the benzoate, as for phenol. The *salicylate ester* [87-18-30] has **m 63-64°** (from aqueous EtOH, or EtOH). [*Beilstein* **6** IV 3296.]

**p-tert-Butylphenoxyacetic acid** [1798-04-5] **M 208.3, m 86.5°, 88-89°, 94°, 96.5°, pK<sub>Est</sub> ~2.9.** Crystallise the acid from petroleum ether or petroleum ether/\*C<sub>6</sub>H<sub>6</sub> mixture. [*Beilstein* **6** H 524, **6** III 1869.]

**tert-Butyl phenyl carbonate** [6627-89-0] **M 194.2, b 74-78°/0.5mm, 83°/0.6mm, d<sub>4</sub><sup>20</sup> 1.05, n<sub>D</sub><sup>20</sup> 1.480.** If IR is free from OH, then purify it by redistillation; otherwise dissolve it in Et<sub>2</sub>O, wash it with 5% HCl, then H<sub>2</sub>O, dry it (MgSO<sub>4</sub>), evaporate and distil it through a Claisen head under vacuum. Care should be taken as distillation of large quantities can lead to decomposition with liberation of CO<sub>2</sub> and isobutylene; **take the necessary precautions**. [Carpino *J Am Chem Soc* **79** 98 1957, *Beilstein* **6** IV 629.]

**n-Butyl phenyl ether** [1126-79-0] **M 150.2, b 95°/17mm, 210.20°/760mm, d<sub>4</sub><sup>20</sup> 0.935, n<sub>D</sub><sup>20</sup> 1.4969.** Dissolve it in diethyl ether, washed first with 10% aqueous NaOH to remove traces of phenol, then repeatedly with distilled water, followed by evaporation of the solvent and distillation under reduced pressure [Arnett & Wu *J Am Chem Soc* **82** 5660 1960]. [*Beilstein* **6** H 143, **6** I 82, **6** II 145, **6** III 550, **6** IV 558.]

**N-tert-Butyl α-phenyl nitron (PBN)** [3376-24-7] **M 177.2, m 73-74°.** Crystallise PBN from hexane. It is a free radical trap. [*cf* Janzen *Methods Enzymology* **105** 188 1984, *Beilstein* **7** IV 519.]

**p-tert-Butyltoluene** [98-51-1] **M 148.3, f -53.2°, b 91°/28mm, 189-192°/atm, d<sub>4</sub><sup>20</sup> 0.854, n<sub>D</sub><sup>20</sup> 1.4920.** A sample containing 5% of the *meta*-isomer is purified by selective mercuration. Fractional distillation of the solid arylmercuric acetate, after removal from the residual hydrocarbon, gives pure *p-tert*-butyltoluene [Stock & Brown *J Am Chem Soc* **81** 5615 1959]. [*Beilstein* **5** H 439, **5** III 1003, **5** IV 1079.]

**tert-Butyl 2,4,5-trichlorophenyl carbonate** [16965-08-5] **M 297.6, m 64-66°, 67-68.5°.** Crystallise the carbonate from a mixture of MeOH (90ml) and water (6ml) using charcoal [Broadbent et al. *J Chem Soc (C)* 2632 1967, Fieser & Fieser *Reagents for Organic Synthesis* **2** 55 1969].

**Caffeic acid (3,4-dihydroxycinnamic acid)** [331-39-5] **M 180.2, m 195°, 223-225°, pK<sub>1</sub><sup>25</sup> 4.62, pK<sub>2</sub><sup>25</sup> 9.07.**

Recrystallise this antioxidant from water. [*Beilstein* 10 IV 1776.]

**Calcon carboxylic acid** [3-hydroxy-4-(2-hydroxy-4-sulfo-1-naphthylazo)naphthalene-2-carboxylic acid] [3737-95-9] **M 428.4, m 300°**,  $\lambda_{\max}$  560nm, **pK<sub>1</sub> 1.2, pK<sub>2</sub> 3.8, pK<sub>3</sub> 9.26, pK<sub>4</sub> 13.14**. Purify it via its *p*-toluidinium salt, viz: dissolve the dye in warm 20% aqueous MeOH, treat with an equivalent of *p*-toluidine and cool to precipitate the salt. Finally recrystallise the acid from hot water. [Itoh & Ueno *Analyst (London)* 95 583 1970.] It is an indicator and complexes with Ca<sup>2+</sup> in presence of Mg<sup>2+</sup> and other metal ions [Patton & Reeder *Anal Chem* 28 1026 1956, Prentoe & Prentoe *Analyst* 106 227 1981].

**Calmagite** [1-(1-hydroxy-4-methyl-2-phenylazo)-2-hydroxynaphthalene-4-sulfonic acid] [3147-14-6] **M 358.4, m 300°**, **pK<sub>1</sub> 8.1, pK<sub>2</sub> 12.4**. A crude dye is extracted with anhydrous diethyl ether and forms red crystals from Me<sub>2</sub>CO. It gives a red colour in H<sub>2</sub>O at pH 7—9 and a blue colour at pH 9—11 which turns red on addition of Ca<sup>2+</sup> or Mg<sup>2+</sup> ions. [Lindstrom & Diehl *Anal Chem* 32 1123 1960]. Complexes with Ca, Mg and Th.

**Capsaicin** (*E-N*-[(4-hydroxy-3-methoxyphenyl)-methyl]-8-methyl-6-nonenamide) [404-86-4] **M 305.4, m 64-66°, 65°, 66.1°, b 210-220°/0.01mm,  $\epsilon$  7000 (281nm, EtOH)**. Recrystallise capsaicin from petroleum ether (b 40-60°), or petroleum ether/Et<sub>2</sub>O (9:1). Also purify it by chromatography on neutral Al<sub>2</sub>O<sub>3</sub> (grade V) and elute successively with \*C<sub>6</sub>H<sub>6</sub>, \*C<sub>6</sub>H<sub>6</sub>/EtOAc (17:3) then \*C<sub>6</sub>H<sub>6</sub>/EtOAc (7:3), and distil it at 120°/10<sup>-5</sup>mm, then repeatedly recrystallise the needles from isopropanol (charcoal). [Crombie et al. *J Chem Soc* 11025 1955, Bennett & Kirby *J Chem Soc(C)* 442 1968.] It **causes pain** and is **neurotoxic** [Bevan & Szolcsanyi *Trends in Pharmacol Sci* 11 330 1990, *Beilstein* 13 IV 2588].

**4-(Carbamoylmethoxy)acetanilide** [14260-41-4] **M 208.2, m 208°**. Crystallise the anilide from water.

***N*-Carboethoxyphthalimide** (*N*-ethoxycarbonylphthalimide) [22509-74-6] **M 219.2, m 87-89°, 90-92°**. Crystallise the imide from toluene/petroleum ether (or \*benzene/petroleum ether). It is partly soluble in Et<sub>2</sub>O, \*benzene and CHCl<sub>3</sub>. [Heller & Jacobsohn *Chem Ber* 54 1112 1921, *Beilstein* 21/10 V 428.]

**3-Carboxybenzaldehyde** (3-formylbenzoic acid, isophthalaldehydic acid) [619-21-6] **M 150.1, m 173°, 173-175°, pK<sup>25</sup> 3.84**. The acid was prepared in several ways including the following two. *m*-Bromomethylbenzoic acid (3g) is allowed to react with hexamine (3.9g) in CHCl<sub>3</sub> (30ml) by refluxing for 3 hours, cooled, the hexamine quaternary salt (4.7g) is collected, dried *in vacuo*, and decomposed by boiling in 50% aqueous acetic acid for 2 hours. The solution is cooled, diluted with an equal volume of H<sub>2</sub>O and chilled to give a crop of the desired acid as needles, which are washed with aqueous NaHCO<sub>3</sub> and dried. A further crop can be obtained by extracting the acidic filtrate with Et<sub>2</sub>O to give total yield of 48%. [Dyer et al. *J Chem Soc* 4778, 4781 1952]. Alternatively, *m*-carboxycinnamic acid (m 268-270°, 8.6g) in 1N Na<sub>2</sub>CO<sub>3</sub> (50ml) and H<sub>2</sub>O (400ml) is treated slowly at 4-5° with 5% of aqueous KMnO<sub>4</sub> (80ml) with vigorous stirring. MnO<sub>2</sub> is removed by filtration, the filtrate is concentrated to 150ml, acidified with 5N HCl and the mixture of acids that separated, together with the solution, are extracted thoroughly with Et<sub>2</sub>O, which on evaporation give 3-formylbenzoic acid (3.6g) which is purified by recrystallisation from H<sub>2</sub>O. [Berner *Acta Chim Scand* 10 1208 1956].

The *acid chloride* [75650-38-3] has **b 130°/20mm**, the *methyl ester* [52178-50-4] has **m 53°**, the *ethyl ester* [33745-47-0] has **d<sup>18</sup> 1.093** and **b 278°/atm**, the *amide* has **m 190°(dec)**, the *oxime* has **m 188°**, the *semicarbazone* has **m 275°(dec)** (from aqueous EtOH), the *bis(2-hydroxyethyl)amine salt* has **m 167°** (yellow crystals from aqueous EtOH), and the *phenylhydrazone* has **m 164°**. [Irreverre et al. *J Biol Chem* 236 1093 1961, Davies et al. *J Chem Soc* 2202 1922, Simonis *Chem Ber* 45 1584 1912, *Beilstein* 10 H 671, 10 I 317, 10 II 465, 10 III 2990, 10 IV 2750.]

**4-Carboxybenzaldehyde** (4-formylbenzoic acid, terephthalaldehydic acid) [619-66-9] **M 150.1, m 247°, 248°(dec), ~250°, 248-250°, 256°, pK<sup>25</sup> 3.77**. Of the many syntheses of this acid, the more convenient one is by refluxing a solution of the commercially available 4-bromomethylbenzoic acid (21.0g, 6232-88-8) with 10% Cu(NO<sub>2</sub>)<sub>2</sub> (300ml) for 6 hours, cooling, filtering off the solid and recrystallising it from glacial AcOH (100ml)

to give fine plates, which on further recrystallisation from H<sub>2</sub>O, provide pure *4-formylbenzoic acid* as fine needles (2.2g 56%). The *phenylhydrazone* has **m 221-222°** [Stewart & Walker *Can J Chem* **35** 1561 1957], the *2,4-dinitrophenylhydrazone* has **m 319.5-320.5°**, after crystallisation from pyridine then PhNO<sub>2</sub> [Bowen & Wilkinson *J Chem Soc* 750 1950], the *oxime* has **m 216-217°** [Wheeler et al. *J Org Chem* **22** 547 1957]. The acid is also obtained by hydrolysis of the *methyl ester*, **m 61-61°** (from petroleum ether), **b 135°/12mm**, by aqueous H<sub>2</sub>SO<sub>4</sub>; the ester having been prepared from methyl 4-cyanobenzoate by reaction with ZnCl<sub>2</sub>/HCl [Slota & Kethur *Chem Ber* **71** 335 1938]. The UV in hexane of the *formyl acid* has  $\lambda_{\max}$  ( $\epsilon$ ) at 249 (17,500), 257 (15,500) [*B-band*], and 279 (1,650), 288 (1,950), 298 (1,600) [*C-band*] nm [Dearden & Forbes *Can J Chem* **36** 1362 1958]. [for pK see Humfray et al. *JCS, Chem Comm* 610 1965; *Beilstein* **10** H 671, **10** I 317, **10** II 465, **10** III 2989, **10** IV 2752.]

**4-Carboxyphenylacetonitrile** [6627-91-4] **M 161.2, m 114-115°**. Crystallise the nitrile (with considerable loss) from \*benzene, glacial acetic acid or H<sub>2</sub>O. The *methyl ester* has **m 47-48°** (from \*C<sub>6</sub>H<sub>6</sub>). [Price & Rogers *Org Synth Coll Vol III* 174 1955, *Beilstein* **9** H 859, **9** II 618, **9** III 4267.]

**Catechol (1,2-dihydroxybenzene, pyrocatechol)** [120-80-9] **M 110.1, m 105°, pK<sub>1</sub><sup>25</sup> 9.45, pK<sub>2</sub><sup>25</sup> 12.8**. Crystallise catechol from \*benzene or toluene and sublime it *in vacuo*. [Rozo et al. *Anal Chem* **58** 2988 1986, *Beilstein* **6** IV 5557.]

**Cation exchange resin**. The resin should be conditioned before use by successive washing with water, EtOH and water, and taken through two H<sup>+</sup>-Na<sup>+</sup>-H<sup>+</sup> cycles by successive treatment with M NaOH, water and M HCl then washed with water until neutral.

***p*-Chloranil (2,3,5,6-tetrachloro-1,4-benzoquinone)** [118-75-2] **M 245.9, m 290°, 294.2-294.6°(sealed tube)**. Crystallise *p*-chloranil from acetic acid, acetone, \*benzene, EtOH or toluene, dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>, or from acetic acid and drying over NaOH in a vacuum desiccator. It can be sublimed under vacuum at 290°. A sample may contain significant amounts of the *o*-chloranil isomer as impurity. Purify it by triple sublimation under vacuum and recrystallise before use. **It is a skin and mucous membrane irritant**. [UV: Pummerer et al. *Chem Ber* **85** 545 1952, Brook *J Chem Soc* 5040 1952, *Beilstein* **7** IV 2083.]

**Chloranilic acid (2,5-dichloro-3,6-dihydroxy-1,4-benzoquinone)** [87-88-7] **M 209.0, m 283-284° pK<sub>1</sub><sup>25</sup> 1.22, pK<sub>2</sub><sup>25</sup> 3.01**. A solution of 8g of quinone in 1L of boiling water is filtered while hot, then extracted twice at about 50° with 200ml portions of \*benzene. The aqueous phase is cooled in ice-water. The crystals are filtered off, washed with three 10ml portions of water, and dried at 115°. It can be sublimed *in vacuo*. [Weissbart & Rysseberghe *J Phys Chem* **61** 765 1957.] The *diacetate* has **m 182-185°** [Conant & Fieser *J Am Chem Soc* **46** 1866 1924, Thamer & Voight *J Phys Chem* **56** 225 1952]. [*Beilstein* **8** IV 2707.]

***p*-Chloroacetanilide** [539-03-7] **M 169.6, m 179°**. Crystallise the anilide from EtOH or aqueous EtOH. [*Beilstein* **12** IV 1178.]

**2-Chloroacetophenone** [532-27-4] **M 154.6, m 54-56°**. Crystallise it from MeOH [Tanner *J Org Chem* **52** 2142 1987]. [*Beilstein* **7** IV 641.]

***o*-Chloroaniline** [95-51-2] **M 127.6, m -1.9°, b 84.5°/10mm, 108.4°/30mm, 208.8°/760mm, d<sub>4</sub><sup>20</sup> 1.213, n<sub>D</sub><sup>20</sup> 1.588, pK<sup>25</sup> 2.66**. Free it from small amounts of the *p*-isomer by dissolving in one equivalent of H<sub>2</sub>SO<sub>4</sub> and steam distilling. The *p*-isomer remains behind as the sulfate. [Sidgwick & Rubie *J Chem Soc* 1013 1921.] An alternative method is to dissolve it in warm 10% HCl (11ml/g of amine) and on cooling, *o*-chloroaniline hydrochloride separates out. The latter can be recrystallised until the acetyl derivative has a constant melting point (**m 90°**). (In this way, yields are better than *via* the recrystallisation of the *picrate* from EtOH or of the acetyl derivative from petroleum ether.) [King & Orton *J Chem Soc* 1377 1911]. [*Beilstein* **12** III 1281, **12** IV 1115.]

***p*-Chloroaniline** [106-47-8] **M 127.6, m 70-71°, b 106.8-107.3-.5°/12mm, 116°/17mm, d<sub>4</sub><sup>20</sup> 1.175, pK<sup>25</sup> 3.98**.

Crystallise the aniline from MeOH, petroleum ether (b 30-60°), or 50% aqueous EtOH, then \*benzene/petroleum ether (b 60-70°), and then dry it in a vacuum desiccator. It can be distilled under vacuum (b 75-77°/3mm). It sublimes in a very high vacuum. The *acetate* crystallises from aqueous MeOH (m 178°, 180°) or EtOH or AcOH (m 173-174°) and has b 331.3°/760mm. [*Beilstein* 12 III 1325, 12 IV 1116.]

*p*-Chloroanisole [623-12-1] M 142.6, b 79°/11.5mm, 196.6°/760mm,  $d_4^{20}$  1.164,  $n_D^{25.5}$  1.5326. Wash the anisole with 10% (by volume) aqueous H<sub>2</sub>SO<sub>4</sub> (three times), 10% aqueous KOH (three times), and then with water until neutral. Dry it (MgSO<sub>4</sub>), and fractionally distil it from CaH<sub>2</sub> through a glass helices-packed column under reduced pressure. [*Beilstein* 16 IV 822.]

9-Chloroanthracene [716-53-0] M 212.9, m 104-106°, 105-107°. 9-Chloroanthracene crystallises from EtOH or petroleum ether (b 60-80°) as yellow needles. [*Nonhebel Org Synth Coll Vol V* 206 1973, *Masnori J Am Chem Soc* 108 1126 1986, *Beilstein* 5 H 663, 5 III 2133, 5 IV 2292.]

10-Chloro-9-anthraldehyde [10527-16-9] M 240.7, m 218°, 217-219°. The aldehyde crystallises as yellow needles from EtOH, AcOH or toluene. [*Beilstein* 7 III 2529.]

2-Chlorobenzaldehyde [89-98-5] M 140.6, m 11°, b 213-214°,  $d_4^{20}$  1.248,  $n_D^{20}$  1.566. Wash it with 10% Na<sub>2</sub>CO<sub>3</sub> solution, then fractionally distil it in the presence of a small amount of catechol as stabiliser. [*Beilstein* 7 H 233, 7 IV 561.]

3-Chlorobenzaldehyde [587-04-2] M 140.6, m 18°, b 213-214°,  $d_4^{20}$  1.241,  $n_D^{20}$  1.564. Purify it by low temperature crystallisation from petroleum ether (b 40-60°) and distillation. [*Beilstein* 7 H 234, 7 IV 566.]

4-Chlorobenzaldehyde [104-88-1] M 140.6, m 47°. Crystallise it from EtOH/water (3:1), then sublime it twice at ~50°/2mm. [*Beilstein* 7 H 235, 7 IV 568.]

Chlorobenzene [108-90-7] M 112.6, b 131.7°,  $d_4^{20}$  1.107,  $n_D^{20}$  1.52480. The main impurities are likely to be chlorinated impurities originally present in the \*benzene used in the synthesis of chlorobenzene, and also unchlorinated hydrocarbons. A common purification procedure is to wash it several times with conc H<sub>2</sub>SO<sub>4</sub> then with aqueous NaHCO<sub>3</sub> or Na<sub>2</sub>CO<sub>3</sub>, and water, followed by drying with CaCl<sub>2</sub>, K<sub>2</sub>CO<sub>3</sub> or CaSO<sub>4</sub>, then with P<sub>2</sub>O<sub>5</sub>, and distilling. It can also be dried with Linde 4A molecular sieve. Passage through, and storage over, activated alumina has been used to obtain low conductance material. [*Flaherty & Stern J Am Chem Soc* 80 1034 1958, *Beilstein* 5 H 199, 5 IV 640.]

4-Chlorobenzenesulfonyl chloride [98-60-2] M 211.1, m 53°, b 141°/15mm. Crystallise it from ether in powdered Dry-ice, after the solution has been washed with 10% NaOH until colourless and dried (Na<sub>2</sub>SO<sub>4</sub>). Distil it *in vacuo* and store it in the absence of H<sub>2</sub>O. [*Beilstein* 11 IV 114.]

4-Chlorobenzhydrazide [536-40-3] M 170.6, m 164°. Crystallise it from H<sub>2</sub>O. [*Beilstein* 9 III 1368.]

2-Chlorobenzoic acid [118-91-2] M 156.6, m 139-140°, pK<sup>25</sup> 2.91. Crystallise the acid successively from glacial acetic acid, aqueous EtOH, and petroleum ether (b 60-80°). Other solvents include hot water or toluene (ca 4ml/g). The crude material can be initially purified by dissolving 30g in 100ml of hot water containing 10g of Na<sub>2</sub>CO<sub>3</sub>, boiling with 5g of charcoal for 15 minutes, then filtering and adding 31ml of 1:1 aqueous HCl. The precipitate is washed with a little water and dried at 100°. [*Beilstein* 9 IV 956.]

3-Chlorobenzoic acid [535-80-8] M 156.6, m 154-156°, 158°,  $d_4^{25}$  1.496, pK<sup>25</sup> 3.82 (5.25 in 50% dimethylacetamide). Crystallise the acid successively from glacial acetic acid, aqueous EtOH and petroleum ether (b 60-80°). It also recrystallises from \*C<sub>6</sub>H<sub>6</sub> or Et<sub>2</sub>O/hexane, and sublimes at 55° in a vacuum. [*Anal Chem* 26 726 1954] The *methyl ester* has m 21°, b 231°/760mm. The *S*-benzylisothiuronium salt has m 164-165° (from EtOH) [*Friediger & Pedersen Acta Chem Scand* 9 1425 1955, *Samuel J Chem Soc* 1318 1960]. [*Beilstein* 9 IV 969.]

**4-Chlorobenzoic acid** [74-11-3] **M 156.6, m 238-239°, pK<sup>25</sup> 3.99.** Purify it as for *m*-chlorobenzoic acid. It has also been crystallised from hot water, and from EtOH. [Beilstein 9 IV 973.]

**2-Chlorobenzonitrile** [873-32-5] **M 137.6, m 45-46°.** Crystallise the nitrile to a constant melting point from \*benzene/petroleum ether (b 40-60°). [Beilstein 9 IV 965.]

**4-Chlorobenzophenone** [134-85-0] **M 216.7, m 75-76°.** Recrystallise it from EtOH. [Wagner et al. *J Am Chem Soc* 108 7727 1986, Beilstein 7 H 419, 7 I 227, 7 II 359, 7 III 2072, 7 IV 1375.]

***o*-Chlorobenzotrifluoride (*o*-chlorotrifluoromethylbenene)** [88-16-4] **M 180.6, m -6.37°, b 19.6°/3mm, 152.3°/760mm, d<sub>4</sub><sup>25</sup> 1.364, n<sub>D</sub><sup>25</sup> 1.4533.** Dry the trifluoride over CaSO<sub>4</sub>, and distil it at high reflux ratio through a silvered vacuum-jacketed glass column packed with one-eighth inch glass helices [Potter & Saylor *J Am Chem Soc* 73 90 1951]. [Beilstein 5 H 302, 5 III 692, 5 IV 814.]

***m*-Chlorobenzotrifluoride** [98-15-7] **M 180.6, m -56.49°, b 50°/31mm, 137.6°/760mm, d<sub>4</sub><sup>20</sup> 1.3345, n<sub>D</sub><sup>20</sup> 1.4432.** Purify it as for *o*-chlorobenzotrifluoride above. [Beilstein 5 III 692, 5 IV 814.]

***p*-Chlorobenzotrifluoride** [98-56-6] **M 180.6, m -33.18°, b 19.3°/5mm, 138.6°/760mm, d<sub>4</sub><sup>30</sup> 1.3278, n<sub>D</sub><sup>30</sup> 1.4430.** Purify it as for *o*-chlorobenzotrifluoride above. [Beilstein 5 IV 815.]

***p*-Chlorobenzyl chloride** [104-83-6] **M 161.0, m 28-29°, b 96°/15mm.** Dry it over CaSO<sub>4</sub>, then fractionally distil it under reduced pressure. Crystallise it from heptane or dry diethyl ether at low temperature. [Beilstein 5 IV 816.] **LACHRYMATORY.**

***p*-Chlorobenzylisothiuronium chloride** [544-47-8] **M 237.1, m 177-178°, and 197°, 201-203°, pK<sub>Est</sub> ~9.6 (free base).** Crystallise the salt from conc HCl by addition of water. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. Also crystallise it from EtOH, wash the crystals with EtOH, then Et<sub>2</sub>O to give the lower melting form **m 177-178°.** By evaporating the filtrate and washings to a quarter of the volume and adding an equal volume of Et<sub>2</sub>O the higher melting form **m 201-203°** is obtained. [Harvey & Jensen *J Org Chem* 28 470 1963, Beilstein 6 III 1639, 6 IV 2778.]

***trans*-4-Chlorocinnamic acid** [1615-02-7] **M 182.6, m 243°, 248-250°, 249-251°, pK<sup>25</sup> 4.41.** Recrystallise the acid from EtOH or aqueous EtOH (charcoal). Its UV has λ<sub>max</sub> at 275nm (EtOH). [Walling & Wolfstirn *J Am Chem Soc* 69 852 1947, Beilstein 9 H 596, 9 II 395, 9 III 2727, 9 IV 2033.]

**4-Chloro-3,5-dimethylphenol** [88-04-0] **M 156.6, m 115.5°, pK<sup>25</sup> 9.70.** Crystallise the phenol from \*benzene or toluene. [Beilstein 6 IV 3152.]

**1-Chloro-2,4-dinitrobenzene** [97-00-7] **M 202.6, m 48-50°, 51°, 52-54°, 54°, b 315°/atm, d<sub>4</sub><sup>22</sup> 1.697.** Usually it is recrystallised from EtOH or MeOH. It has also been crystallised from Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub>, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or isopropyl alcohol. A preliminary purification step is to pass its solution in \*benzene through an alumina column. It has also been purified by zone refining. It exists in three forms: one stable and two unstable. The stable form crystallises as yellow needles from Et<sub>2</sub>O, **m 51°, b 315°/760mm** with some decomposition, and is soluble in EtOH. A labile form also crystallises from Et<sub>2</sub>O, **m 43°**, and is more soluble in organic solvents. The second labile form has **m 27°**. [Hoffman & Dame, *J Am Chem Soc* 41 1015 1919, Welsh *J Am Chem Soc* 63 3276 1941, *J Chem Soc* 2476 1957, Beilstein 5 IV 744.]

**4-Chloro-3,5-dinitrobenzoic acid** [118-97-8] **M 246.6, m 159-161°, 163°, pK<sub>Est</sub> ~2.5.** Crystallise the acid from EtOH/H<sub>2</sub>O, EtOH or \*C<sub>6</sub>H<sub>6</sub>. The 1:1 *naphthalene complex* (by fusing various ratios of ingredients and recrystallising from EtOH) has **m 122°**. [Beilstein 9 H 416, 9 III 1953, 9 IV 1360.]

**Chlorogenic [1-(3,4-dihydroxycinnamoyloxy)-D-quinic] acid** [327-97-9] **M 354.3, m 208°, [α]<sub>D</sub><sup>25</sup> -36° (c 1, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 3.59, pK<sub>2</sub><sup>25</sup> 8.59.** Crystallise the acid from water and dry it at 110°. [Beilstein 10 H 537, 10 I 271, 10 II 378, 10 III 2408, 10 IV 2259.]

**Chlorohydroquinone (2-chloro-1,4-dihydroxybenzene)** [615-67-8] **M 144.6, m 106°**, **b 263°**, **pK<sub>i</sub><sup>25</sup> 8.81, pK<sub>2</sub><sup>25</sup> 10.78**. Crystallise the hydroquinone from CHCl<sub>3</sub> or toluene. [Beilstein 6 IV 5767.]

**4-Chloriodobenzene** [637-87-6] **M 238.5, m 53-54°, 56.2°**, **b 104.2°/16mm, d<sub>4</sub><sup>57</sup> 1.886**. Distil it in a vacuum then recrystallise it from EtOH. [Sugden *J Chem Soc* 1173 1924, Beilstein 5 H 221, 5 III 579, 5 IV 695.]

**5-Chloro-2-methoxyaniline (2-amino-4-chloroanisole)** [95-03-4] **M 157.6, m 81-83°, 82-84°, 84°**, **pK<sup>25</sup> 3.56**. Purify the aniline by steam distillation and recrystallisation from H<sub>2</sub>O or 40% aqueous EtOH. The *N*-acetate forms needles from hot H<sub>2</sub>O with **m 104°**, the *N*-benzoyl derivative forms needles from aqueous EtOH with **m 77-78°**, and the *picrate* has **m 194°(dec)**. [Raiford & Colbert *J Am Chem Soc* 48 2657 1926, Beilstein 13 IV 879.]

**3-Chloro-4-methoxyphenethylamine** [7569-87-1] **M 149.6, b 140°/0.6mm, d<sub>4</sub><sup>25</sup> ~1.081, n<sub>D</sub><sup>20</sup> ~1.553, pK<sup>25</sup> ~9.8**. This strong base is prepared by reduction of *3-chloro-4-methoxyphenylacetone nitrile* (176mmol, **b 140-145°/0.1-0.1mm, m 55-56°**, [7569-58-6]) in THF (100ml) with LiAlH<sub>4</sub> (8.0g, 210mmol) suspended in THF (250ml) under reflux for 5 hours. The greenish coloured solution is carefully decomposed with ice cold H<sub>2</sub>O, the solids are filtered off, washed with Et<sub>2</sub>O, the combined organic liquids are dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated and the crude brown residual amine is distilled in as high a vacuum as possible to give a clear oil (11.4g, 35%). It absorbs CO<sub>2</sub> from air and is best stored in an inert atmosphere. Its <sup>1</sup>H NMR [60MHz, CDCl<sub>3</sub>] has δ at 7.20 (d, 1H, ArH2), 7.04 (q, 1H, ArH6), 6.83 (d, 1H, ArH5), 3.88 (s, 3H, OCH<sub>3</sub>), 3.12-2.20 (m, 4H, PhCH<sub>2</sub>CH<sub>2</sub>N), 1.28 (s, 2H, NH<sub>2</sub>) ppm from TMS [Charifson et al. *J Med Chem* 31 1941 1988]. Alternatively, the amine is prepared from the respective phenethyl bromide and dry NH<sub>3</sub> in EtOH followed by dilution with Et<sub>2</sub>O, washing with aqueous NaOH, and evaporation. The residue is dissolved in 5% aqueous HCl and evaporated to dryness to give *3-chloro-4-methoxyphenethylamine hydrochloride* [7569-60-0] **M 186.1, m 192-195°**, as an apparently amorphous white powder when crystallisation from EtOH/Et<sub>2</sub>O is attempted [Fosdick et al. *J Am Chem Soc* 68 840 1946]. The *N*-benzoyl derivative [115514-67-5] **m 137-140°** is a colourless solid which is insoluble in Et<sub>2</sub>O. [Beilstein 13 III 1650.]

**9-Chloromethyl anthracene** [24463-19-2] **M 226.7, m 141-142°(dec), 141-142.5°**. If it is free from OH in the IR then recrystallisation from hexane/\*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub> (as needles). If OH is present, then some solvolysis has occurred. In this case treat 8.5g of it with SOCl<sub>2</sub> (4.8g) in dioxane (60ml) and reflux for 5 hours, then evaporate to dryness and wash the residue with cold \*C<sub>6</sub>H<sub>6</sub> and recrystallise it. With KI/Me<sub>2</sub>CO it forms the *iodomethyl* derivative. [Fierens et al. *Helv Chim Acta* 38 2009 1955, Hunter et al. *J Org Chem* 21 1512 1956, Beilstein 5 III 3152, 5 IV 2313.]

**4-Chloro-2-methylphenol** [1570-64-5] **M 142.6, m 49°, b 112-114°/18mm, 225°/760mm, pK<sup>25</sup> 9.71**. Purify the phenol by crystallisation from petroleum ether (**m 51°**) and by zone melting. [Beilstein 6 H 359, 6 I 174, 6 II 332, 6 III 1264, 6 IV 1987.]

**4-Chloro-3-methylphenol** [59-50-7] **M 142.6, m 66°, b 238°/760mm, pK<sup>25</sup> 9.55**. Crystallise the phenol from petroleum ether or \*C<sub>6</sub>H<sub>6</sub>. [Beilstein 6 H 381, 6 I 187, 6 II 355, 6 III 1315, 6 IV 2064.]

**4-Chloro-2-methylphenoxyacetic acid (MCPA)** [94-74-6] **M 200.6, m 113-117°, 120°, 122-123°, pK<sup>20</sup> 3.62(3.05)**. It is insoluble in H<sub>2</sub>O (solubility is 0.55g/L at 20°) and recrystallises from \*C<sub>6</sub>H<sub>6</sub> or chlorobenzene as plates [Jönsson et al. *Acta Chem Scand* 6 993 1952]. The *S*-benzylisothiuronium salt has **m 164-165°**, and the Cu<sup>2+</sup> salt has **m 247-249°(dec)** [Armarego et al. *Nature* 183 1176 1959, UV: Duvaux & Grabe *Acta Chem Scand* 4 806 1950, IR: Jöberg *Acta Chem Scand* 4 798 1950]. [Beilstein 6 IV 1991.] It is a plant growth substance and a herbicide.

**2-Chloromethyl-2-phenylpropane (neophyl chloride, 1-chloro-2-methyl-2-phenylpropane, β-chloro-tert-butylbenzene)** [515-40-2] **M 168.7, b 53°/1.0mm, 95.1-95.2°/10mm, 97°/13.0mm, 104°/18.0mm, 120°/30mm, 111°/90.0mm, 222°/741mm (dec), d<sub>4</sub><sup>25</sup> 1.5228, n<sub>D</sub><sup>20</sup> 1.5250**. It is prepared by adding β-methyl chloride (603g, 6.66 moles, CH<sub>2</sub>=C(Me)-CH<sub>2</sub>Cl, 3-chloro-2-methyl-1-propene, prepared by the chlorination of butylene, **b 71.5-72.5°/760mm, d<sub>4</sub><sup>20</sup> 0.9165, n<sub>D</sub><sup>20</sup> 1.4274** [563-47-3], Beilstein 1 IV 803) into a vigorously stirred mixture of

\*benzene (1404g, 18.5 moles, washed twice with conc H<sub>2</sub>SO<sub>4</sub> and used as such) and concentrated H<sub>2</sub>SO<sub>4</sub> (104g, 1 mole) at 20° which required 12 hours, and stirring is continued for a further 11 hours at room temperature. The organic layer is collected, excess of \*C<sub>6</sub>H<sub>6</sub> is distilled off and the residue is fractionated through an 8-plate column to give pure (99.1 ± 0.3% by acetolysis) *neophyl chloride* (765.5g, 68%). It is less reactive than neopentyl chloride towards Na metal and less reactive still towards EtNa, and both are inert towards most basic reagents. [Whitmore et al. *J Am Chem Soc* **65** 1469 1943, Smith & Sellas *Org Synth Coll Vol IV* 702 1963, *Beilstein* **5** IV 1048.]

It readily forms the Grignard reagent *neophyl magnesium chloride* [35293-35-7] **M 193.0**, with Mg in Et<sub>2</sub>O; and a 0.5M solution of this reagent in Et<sub>2</sub>O is available commercially. It reacts with solid CO<sub>2</sub> to give an 82% yield of *β-phenylisovaleric acid* [1010-48-6] (**m 58-59.5°**, from petroleum ether b 60-90°); and oxidation provides a 72% yield of *2-methyl-2-phenylpropan-1-ol* [2173-69-5] (**b 131°/30mm**, **n<sub>D</sub><sup>20</sup> 1.5261**) whose *phenylurethane* has **m 59.5-60.5°**, *α-naphthylurethane* has **m 91.5-92.5°**, and its *p-toluenesulfonate* has **m 74-75°**. [Whitmore et al. *J Am Chem Soc* **65** 1469 1943, Fainberg & Winstein *J Am Chem Soc* **78** 2763 1956, Winstein et al. *J Am Chem Soc* **74** 1113 1952.]

**Chloromethyl phenyl sulfide** [7205-91-6] **M 158.7**, **b 63°/0.1mm**, **98°/12mm**, **113-115°/20mm**, **d<sub>4</sub><sup>20</sup> 1.184**, **n<sub>D</sub><sup>20</sup> 1.5950**. Dissolve the sulfide in CH<sub>2</sub>Cl<sub>2</sub> or CCl<sub>4</sub> and dry it (CaCl<sub>2</sub>), or pass it through a tube of CaCl<sub>2</sub> and distil it using a fractionating column. *Harmful vapours*. It gives the *sulfone* [7205-98-3] (**b 130°/1mm** and **m 53°** from EtOH) [*Beilstein* **6** IV 1507] on oxidation with permonophthalic acid. [Böhme et al. *Justus Liebigs Ann Chem* **563** 54 64 1949.] [*Beilstein* **6** III 1002.]

**N-(Chloromethyl)phthalimide** [17564-64-6] **M 195.6**, **m 131-135°**, **134-135°**, **136.5°**. Purify the imide by recrystallisation from EtOAc or CCl<sub>4</sub> or *via* the 1:1 complex with pyridine [Sakellarios *J Am Chem Soc* **70** 2822 1948, Böhme et al. *Chem Ber* **92** 1258 1959]. [*Beilstein* **21/10** V 372.]

**1-Chloronaphthalene** [90-13-1] **M 162.6**, **f -2.3°**, **b 136-136.5°/20mm**, **259.3°/760mm**, **d<sub>4</sub><sup>20</sup> 1.194**, **n<sub>D</sub><sup>20</sup> 1.6326**. Wash the chloronaphthalene with dilute NaHCO<sub>3</sub>, then dry it with Na<sub>2</sub>SO<sub>4</sub> and fractionally distil it *in vacuo*. Alternatively, before distillation, it is passed through a column of activated alumina, or dried with CaCl<sub>2</sub>, then distilled from sodium. It can be further purified by fractional crystallisation by partial freezing or by crystallisation of its *picrate* to constant melting point (**m 132-133°**) from EtOH, and recovering it from the *picrate*. [*Beilstein* **5** H 541, **5** III 1570, **5** IV 1658.]

**2-Chloronaphthalene** [91-58-7] **M 162.6**, **m 59.5-60°**, **61°**, **b 121-122°/12mm**, **264-266°/760mm**. Distil 2-chloronaphthalene in a vacuum, then crystallise it from 25% EtOH/water, then dry it under vacuum (see also the 1-isomer above). [*Beilstein* **5** H 541, **5** I 262, **5** II 445, **5** III 1573, **5** IV 1660.]

**1-Chloro-2-naphthol** [633-99-8] **M 178.6**, **m 70°**, **71°**, **pK<sub>Est</sub> ~8.3**. Crystallise the naphthol from petroleum ether. The *acetate* has **m 42-43°**. [*Beilstein* **6** I 315, **6** II 603, **6** III 2990, **6** IV 4289.]

**2-Chloro-1-naphthol** [606-40-6] **M 178.6**, **m 64-65°**, **65°**, **pK<sup>20</sup> 9.9 (aqueous EtOH)**. Crystallise the naphthol from petroleum ether. [*Beilstein* **6** I 308, **6** II 581, **6** III 2933, **6** IV 4230.]

**4-Chloro-1-naphthol** [604-44-4] **M 178.6**, **m 116-117°**, **120-121°**, **pK<sup>25</sup> 8.86, 10.7 (aqueous EtOH)**. Crystallise the naphthol from EtOH or CHCl<sub>3</sub>. [*Beilstein* **6** H 611, **6** II 582, **6** III 2933, **6** IV 4233.]

**4-Chloro-2-nitroaniline** [89-63-4] **M 172.6**, **m 114-115°**, **116-116.5°**, **pK<sup>25</sup> -0.99**. Crystallise the aniline from hot H<sub>2</sub>O (**m 115.8-116°**), EtOH, EtOH/H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>, and dry it for 10 hours at 60° *in vacuo*. It has **m 115.5-116°** after sublimation. [*Beilstein* **12** H 729, **12** I 355, **12** II 396, **12** III 1649, **12** IV 1669.]

**2-Chloro-4-nitrobenzamide** [3011-89-0] **M 200.6**, **m 170-171°**, **172°**. Crystallise the amide from EtOH. [Jensen & Ploug *Acta Chem Acta* **3** 15 1949, *Beilstein* **9** H 404, **9** III 1768.]

**2-Chloro-1-nitrobenzene** [88-73-3] **M 157.6**, **m 32.8-33.2°**. Crystallise it from EtOH, MeOH or pentane (charcoal). [*Beilstein* **5** IV 721.]

**3-Chloro-1-nitrobenzene** [121-73-3] **M 157.6, m 45.3-45.8°**. Crystallise the nitrobenzene from MeOH or 95% EtOH (charcoal), then pentane. [Beilstein 5 IV 722.]

**4-Chloro-1-nitrobenzene** [100-00-5] **M 157.6, m 80-83°, 83.5-84°, b 113°/8mm, 242°/atm, d<sub>4</sub><sup>100.5</sup> 1.2914**. Crystallise the nitrobenzene from 95% EtOH (charcoal) and sublime it *in vacuo*. [Emmons *J Am Chem Soc* 76 3470 1954, Newman & Forrest *J Am Chem Soc* 69 1221 1947, Beilstein 5 IV 723.]

**3-Chloroperbenzoic acid (MCPBA)** [937-14-4] **M 172.6, m 92-94°(dec), pK<sup>25</sup> 7.57**. Recrystallise MCPBA from CH<sub>2</sub>Cl<sub>2</sub> [Traylor & Mikztal *J Am Chem Soc* 109 2770 1987]. Peracid of 99+% purity can be obtained by washing commercial 85% material with phosphate buffer pH 7.5 and drying the residue under reduced pressure. Alternatively, the peracid can be freed from *m*-chlorobenzoic acid by dissolving 50g/L of \*benzene and washing with an aqueous solution buffered at pH 7.4 (NaH<sub>2</sub>PO<sub>4</sub>/NaOH) (5 x 100ml). The organic layer is dried over MgSO<sub>4</sub> and carefully evaporated under vacuum. *Necessary care should be taken in case of EXPLOSION*. The solid is recrystallised twice from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O and stored at 0° in a plastic container as glass catalyses the decomposition of the peracid. The acid is assayed iodometrically. [Schwartz & Blumbrgs *J Org Chem* 29 1976 1964, Bortolini et al. *J Org Chem* 52 5093 1987, McDonald et al. *Org Synth Coll Vol VI* 276 1988, Beilstein 9 IV 972.]

**2-Chlorophenol** [95-57-8] **M 128.6, m 8.8°, b 61-62°/10mm, 176°/760mm, pK<sup>25</sup> 8.34**. Pass 2-chlorophenol at least twice through a gas chromatography column. It has also been purified by fractional distillation. [Beilstein 6 IV 782.]

**3-Chlorophenol** [108-43-0] **M 128.6, m 33°, b 44.2°/1mm, 214°/760mm, pK<sup>25</sup> 9.13**. It could not be obtained solid by crystallisation from petroleum ether. It is best purified by distillation under reduced pressure. [Beilstein 6 IV 810.]

**4-Chlorophenol** [106-48-9] **M 128.6, m 43°, 100-101°/10mm, pK<sup>25</sup> 9.38**. Distil the phenol, then crystallise it from petroleum ether (b 40-60°) or hexane, and dry it under vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. [Bernasconi & Paschalis *J Am Chem Soc* 108 2969 1986, Beilstein 6 IV 820.]

**Chlorophenol Red (3,3'-dichlorophenolsulfonephthalein)** [4430-20-0] **M 423.3, m dec on heating, λ<sub>max</sub> 573nm, pK<sup>25</sup> 5.96**. Crystallise the dye from glacial acetic acid. It is yellow at pH 4.8 and violet at pH 6.7. [Beilstein 19/3 V 458.]

**4-Chlorophenoxyacetic acid** [122-88-3] **M 186.6, m 157°, pK<sup>20</sup> 3.00, 4.15 (50% aqueous EtOH)**. Crystallise the acid from EtOH. It is a plant growth substance and a herbicide. [Beilstein 6 IV 845.]

**(±)-α-4-Chlorophenoxypropionic acid** [3307-39-9] **M 200.6, m 116°, pK<sub>Est</sub> ~3.2**. Crystallise the acid from EtOH or HCOOH (m 114.5-115.5°). It is a plant growth substance. The *R*(+)- and *S*(-)-enantiomers have m 103-104° (from petroleum ether) and [α]<sub>D</sub><sup>25</sup> ±41° (c 1, EtOH). [Beilstein 6 III 695, 6 IV 850.]

**β-4-Chlorophenoxypropionic acid** [3284-79-5] **M 200.6, m 138°, pK<sub>Est</sub> ~4.2**. Crystallise the acid from EtOH. It is a plant growth substance. [Beilstein 6 III 696, 6 IV 851.]

**3-Chlorophenylacetic acid** [1878-65-5] **M 170.6, m 74°, pK<sup>25</sup> 4.11**. Crystallise the acid from EtOH/water, or as needles from \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O (charcoal). The *acid chloride* (prepared by boiling with SOCl<sub>2</sub>) has b 127-129°/15mm. [Dippy & Williams *J Chem Soc* 161 1934, Misra & Shukla *J Indian Chem Soc* 28 480 1951, Beilstein 9 III 2263, 9 IV 1674.]

**4-Chlorophenylacetic acid** [1878-66-6] **M 170.6, m 105°, 106°, pK<sup>25</sup> 4.12**. Purify it as for 3-chlorophenylacetic acid. [Beilstein 9 III 2263, 9 IV 1674.]

**4-Chloro-1-phenylbutan-1-one** [939-52-6] **M 182.7, m 19-20°, b 134-137°/5mm, d<sub>4</sub><sup>20</sup> 1.149, n<sub>D</sub><sup>20</sup> 1.55413**. Fractionate the ketone several times using a short column. It recrystallises from petroleum ether at -20° in

glistening white rosettes and is filtered at 0°, and dried in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. The *semicarbazone* has *m* 136-137°. [Conant et al. *J Am Chem Soc* **46** 1882 1924, Cloke *J Chem Soc* 1174 1929, Hart & Curtis *J Am Chem Soc* **79** 931 1957, *Beilstein* 7 IV 711.]

**1-(2-Chlorophenyl)-1-(4-chlorophenyl)-2,2-dichloroethane (Mitotane, *o,p'*-DDD)** [53-19-0] *M* 320.1, *m* 75.8-76.8°, 76-78°. Purify Mitotane by recrystallisation from pentane, MeOH or EtOH. It is soluble in isooctane and CCl<sub>4</sub>. [Haller et al. *J Am Chem Soc* **67** 1600 1945, *Beilstein* 5 IV 1883.]

**3-(4-Chlorophenyl)-1,1-dimethylurea (Monuron)** [150-68-5] *M* 198.7, *m* 171°. Crystallise monuron from MeOH. [*Beilstein* 12 IV 1191.]

**4-Chloro-1,2-phenylenediamine** [95-83-0] *M* 142.6, *m* 69-70°, *pK*<sub>1</sub><sup>25</sup> -0.27 (aqueous H<sub>2</sub>SO<sub>4</sub>), *pK*<sub>2</sub><sup>25</sup> 3.35 (3.67). Recrystallise the diamine from petroleum ether. [*Beilstein* 13 IV 68.]

**4-Chlorophenyl isocyanate** [104-12-1] *M* 153.6, *m* 28-31°, 31-32°, 32°, 32.5°, *b* 80.6-80.9°/9.5mm, 115-117°/45mm. Purify the isocyanate by recrystallisation from petroleum ether (*b* 30-40°) or better by fractional distillation. **TOXIC irritant.** [*Beilstein* 12 H 616, 12 III 1376, 12 IV 1213.]

**4-Chlorophenyl isothiocyanate** [2131-55-7] *M* 169.6, *m* 44°, 43-45°, 45°, 46°, 47°, *b* 110-115°/4mm, 135-136°/24mm. Check the IR first to see if free from OH frequencies. Triturate it with petroleum ether (*b* 30-60°) and decant the solvent. Repeat this 5 times. The combined extracts are evaporated under reduced pressure to give almost pure compound as a readily crystallisable oil with a pleasant anise odour. It can be recrystallised from the minimum volume of EtOH at 50° (do not boil too long as it could react). It can be purified by vacuum distillation. [van der Kerk et al. *Org Synth Coll Vol V* 223 1973, *Beilstein* 12 IV 1214.] **It is an IRRITANT and causes dermatitis; use gloves.**

**4-Chlorophenyl 2-nitrobenzyl ether** [109669-56-9] *M* 263.7, *m* 44.5°, *b* 154-156°/3mm, 208°/11mm. Distil it under reduced pressure, and it crystallises from EtOH (*m* 44-45°) or MeOH (*m* 46°) as yellow needles. [*Beilstein* 6 II 210, 6 III 801, 6 IV 1253.]

**4-Chlorophenyl 4-nitrobenzyl ether** [5442-44-4] *M* 263.7, *m* 77°, *b* 215°/12mm. Distil it in a vacuum and crystallise it from EtOH, MeOH (*m* 75.5-76°) or petroleum ether (*m* 76°, 77°). Its UV has λ<sub>max</sub> at 222 and 302nm (EtOH). [*Beilstein* 6 II 222, 6 III 821, 6 IV 1288.]

**4-Chlororesorcinol** [95-88-5] *M* 144.6, *m* 105°, *pK*<sub>Est(1)</sub>~9.2, *pK*<sub>Est(2)</sub>~10.1. Crystallise it from boiling CCl<sub>4</sub> (10g/L, charcoal) and dry it in air. [*Beilstein* 6 II 818.] **IRRITANT.**

**5-Chlorosalicylaldehyde** [635-93-8] *M* 156.6, *m* 98.5-99°, 99.5°, 101°, *pK*<sup>25</sup> 7.4. Steam distil it, then crystallise it from aqueous EtOH or \*C<sub>6</sub>H<sub>6</sub> (*m* 100°). It forms complexes with Cu<sup>2+</sup> and Fe<sup>2+</sup>. [*Beilstein* 8 H 53, 8 II 45, 8 III 181, 8 IV 224.]

**4-Chlorothiophenol** [106-54-7] *M* 144.6, *m* 51-52°, 53.5-54°, *pK*<sup>25</sup> 6.14. Recrystallise the thiophenol from aqueous EtOH. The *SMe ether* has *m* 129° and the *SEt ether* has *m* 64°. [D'Sousa et al. *J Org Chem* **52** 1720 1987, *Beilstein* 6 H 326, 6 I 149, 6 III 1034.]

**2-Chlorotoluene** [95-49-8] *M* 126.6, *b* 159°, *d*<sub>4</sub><sup>20</sup> 1.083, *n*<sub>D</sub><sup>20</sup> 1.5255. Dry 2-chlorotoluene for several days, with CaCl<sub>2</sub>, then distil it from Na using a glass helices-packed column. [*Beilstein* 5 IV 805.]

**3-Chlorotoluene** [108-41-8] *M* 126.6, *m* -48°, *b* 161-163°, *d*<sub>4</sub><sup>20</sup> 1.072, *n*<sub>D</sub><sup>20</sup> 1.522. Purify it as for 2-chlorotoluene above. [*Beilstein* 5 IV 806.]

**4-Chlorotoluene** [106-43-4] *M* 126.6, *m* 7.2°, *b* 162.4°, *d*<sub>4</sub><sup>20</sup> 1.07, *n*<sub>D</sub><sup>20</sup> 1.521. Dry it with BaO, fractionally distil it, then fractionally crystallise it by partial freezing. [*Beilstein* 5 IV 806.]

**Chrysene** [218-01-9] **M 228.3, m 255-256°**. Purify chrysene by chromatography on alumina from petroleum ether in a darkened room. Its solution in \*C<sub>6</sub>H<sub>6</sub> is passed through a column of decolorising charcoal, then crystallised by concentrating the eluate. It has also been purified by crystallising from \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, and by zone refining. [Gorman et al. *J Am Chem Soc* **107** 4404 1985]. It is freed from 5*H*-benzo[*b*]carbazole by dissolving it in *N,N*-dimethylformamide and successively adding small portions of alkali and iodomethane until the fluorescent colour of the carbazole anion no longer appears when alkali is added. The chrysene (and alkylated 5*H*-benzo[*b*]carbazole) separate on addition of water. Final purification is by crystallisation from ethylcyclohexane and/or from 2-methoxyethanol [Bender et al. *Anal Chem* **36** 1011 1964]. It can be sublimed in a vacuum. [Beilstein **5** IV 2554.]

**Chrysoidine G (4-phenylazo-1,3-benzenediamine monohydrochloride, basic orange 2)** [532-82-1] **M 248.7, m 118-118.5°, pK<sub>1</sub> 3.32, pK<sub>2</sub> 5.21**. It is a red-brown powder which is recrystallised from H<sub>2</sub>O. It gives a yellow solution in conc H<sub>2</sub>SO<sub>4</sub> which turns orange on dilution. Its solubility at 15° is 5.5% (H<sub>2</sub>O), 4.75% (EtOH), 6.0% (cellosolve), 9.5% (ethylene glycol), 0.005% (xylene) and is insoluble in \*C<sub>6</sub>H<sub>6</sub>. The *hydroiodide* has **m** 184° (from EtOH) and the *picrate* forms red needles with **m** 196°. [Muramatsu *Bull Chem Soc Jpn* **31** 864 1958, *Beilstein* **6** IV 561.]

**trans-Cinnamaldehyde** [14271-10-9] **M 132.2, m -4°, -7.5°, -9°, b 80°/0.4mm, 85.8°/1.1mm, 125-128°/11mm, 152.2°/40mm, 163.7°/60mm, 199.3°/200mm, 246°/760mm (dec), d<sub>4</sub><sup>20</sup> 1.0510, n<sub>D</sub><sup>20</sup> 1.623**. Purify the aldehyde by steam distillation (solubility is 1 in 700 parts H<sub>2</sub>O) followed by distillation *in vacuo*. The *cis*-isomer has **b** 67-69°/40mm and d<sub>4</sub><sup>20</sup> 1.0436 and n<sub>D</sub><sup>20</sup> 1.5937. The *trans-semicarbazone* has **m** 210°(dec) from CHCl<sub>3</sub>/MeOH (*cis-semicarbazone* has **m** 196°), the *trans-phenylsemicarbazone* has **m** 177° from CHCl<sub>3</sub>/MeOH (the *cis-phenylsemicarbazone* has **m** 146°), the *trans-2,4-dinitrophenylhydrazone* has **m** 250°(dec) from MeOH as the *cis*-isomer [Gamboni et al. *Helv Chim Acta* **38** 255 1955, Holum *J Org Chem* **26** 4814 1961]. [*Beilstein* **9** IV 984.]

**cis-Cinnamic acid (Z-3-phenyl-2-propenoic acid)** [102-94-3] **M 148.2, m 68° (for allo-form), pK<sup>25</sup> 3.93**. The *cis*-acid is prepared by catalytic reduction of phenylpropionic acid and after distillation in a high vacuum at ~95° it gives the most stable *allo*-isomer **m** 68°. Recrystallisation from petroleum ether yields **Liebermann's iso-cinnamic acid m 58°**. When the *allo*-acid (**m** 68°) is heated at 20° above its melting point in a sealed capillary for 0.5 hours and allowed to cool slowly, **Erlenmyer's iso-cinnamic acid m 42°** is formed. This form can also be obtained in larger amounts by heating the *allo*-acid at 80° for 3 hours, and on cooling it remains liquid for several weeks but gives the **m** 42° acid on inoculation with the crystals from the capillary tube. This form is unchanged in 6 weeks when kept in a dark cupboard. All three forms have the same pK values and the same rate of bromination. There is also a very labile form with **m** 32°. [Liebermann, *Chem Ber* **26** 1572 1893, Claisen & Crismer *Justus Liebigs Ann Chem* **218** 135 1883, Robinson & James *J Chem Soc* 1453 1933, Berthoud & Urech *Helv Chim Acta* **13** 437 1930, McCoy & McCoy *J Org Chem* **33** 2354 1968, *Beilstein* **9** IV 2001.]

**trans-Cinnamic (E-3-phenyl-2-propenoic) acid** [140-10-3, 621-82-9 for *E-Z* mixture] **M 148.2, m 134.5-135°, pK<sup>25</sup> 4.42 (4.50)**. Crystallise the acid from \*benzene, CCl<sub>4</sub>, hot water, water/EtOH (3:1), or 20% aqueous EtOH. Dry it at 60° *in vacuo*. It is steam volatile. [*Beilstein* **9** IV 2002.]

**trans-Cinnamic anhydride** [538-56-7] **M 278.4, m 136°**. Crystallise the anhydride from \*C<sub>6</sub>H<sub>6</sub> or toluene/petroleum ether (**b** 60-80°) or EtOH (**m** 135-136°). [*Beilstein* **9** III 2703, **9** IV 2018.]

**trans-Cinnamoyl chloride** [102-92-1] **M 166.6, m 35-37°, b 101°/2mm, 154°/25mm, 256-258°/atm, d<sub>4</sub><sup>37.6</sup> 1.6202, n<sub>D</sub><sup>37.6</sup> 1.1632**. Refractionate it in a vacuum until the distillate solidifies on cooling, and recrystallise it from petroleum ether. The *trans-amide* has **m** 145-150° (from H<sub>2</sub>O) [*Beilstein* **9** III 2711]. [Adams & Ulich *J Am Chem Soc* **42** 605 1920, Bergmann et al. *J Chem Soc* 2524 1952, *Beilstein* **9** H 587, **9** I 233, **9** II 390, **9** III 2710, **9** IV 2020.]

***N*-Cinnamoyl-*N*-phenylhydroxylamine** [7369-44-0] **M 239.3, m 158-163°**. Recrystallise the hydroxylamine from EtOH.

**Cinnamyl alcohol** [104-54-1] **M 134.2, m 33°, b 143.5°/14mm,  $\lambda_{\max}$  251nm ( $\epsilon$  18,180 M<sup>-1</sup> cm<sup>-1</sup>).** Crystallise the alcohol from diethyl ether/pentane. [Beilstein 6 I 281.]

**Congo Red (4B) (cotton red B)** [573-58-0] **M 696.7, m >360°,  $\lambda_{\max}$  497nm, pK<sub>2</sub><sup>28</sup> 4.19.** Crystallise the dye from aqueous EtOH (1:3). Dry it in air. [Beilstein 6 I 342.]

**Coniferyl alcohol** [4-hydroxy-3-methoxy-cinnamyl alcohol, 3-(4-hydroxy-3-methoxy-phenyl)-2-propen-1-ol] [458-35-5] **M 180.2, m 73-75°, b 163-165°/3mm, pK<sup>25</sup> 9.54.** It is soluble in EtOH and insoluble in H<sub>2</sub>O. It can, however, be recrystallised from EtOH and distilled in a vacuum. It polymerises in dilute acid. The *benzoyl* derivative has **m 95-96°** (from petroleum ether), and the *tosylate* has **m 66°**. [Derivatives: Freudenberg & Achtzehn *Chem Ber* 88 10 1955, UV: Herzog & Hillmer *Chem Ber* 64 1288 1931, Beilstein 6 II 1093.]

**Coronene** [191-07-1] **M 300.4, m 438-440°, 442°, b 525°,  $\lambda_{\max}$  345nm (log  $\epsilon$  4.07).** Crystallise coronene from \*benzene or toluene, then sublime it in a vacuum. [Beilstein 5 III 2651.]

***o*-Cresol** [95-48-7] **M 108.1, m 30.9°, b 191°/760mm, n<sub>D</sub><sup>41</sup> 1.536, n<sub>D</sub><sup>46</sup> 1.534, pK<sup>25</sup> 10.22.** It can be freed from *m*- and *p*-isomers by repeated fractional distillation. It crystallises from \*benzene by addition of petroleum ether. It has been fractionally crystallised by partial freezing of its melt. The *3,5-dinitrobenzoate* (prepared from 3,5-dinitrobenzoyl chloride in dry pyridine, and recrystallised from EtOH or aqueous Me<sub>2</sub>CO) has **m 138°**. [Beilstein 6 IV 1940.]

***m*-Cresol** [108-39-4] **M 108.1, f 12.0°, b 202.7°, d<sub>4</sub><sup>20</sup> 1.034, n<sub>D</sub><sup>20</sup> 1.544, pK<sup>25</sup> 0.09.** Separation of the *m*- and *p*-cresols requires chemical methods, such as conversion to their sulfonates [Brüchner *Anal Chem* 75 289 1928]. An equal volume of H<sub>2</sub>SO<sub>4</sub> is added to *m*-cresol, stirred with a glass rod until solution is complete. Heat for 3 hours at 103-105°. Dilute carefully with 1-1.5 volumes of water, heat to boiling point and steam distil until all unsulfonated cresol has been removed. Cool and extract the residue with ether. Evaporate the solution until the boiling point reaches 134° and steam distil off the *m*-cresol. Another purification method involves distillation, fractional crystallisation from the melt, then redistillation. Free from *p*-cresol by solution in glacial acetic acid and bromination by about half of an equivalent amount of bromine in glacial acetic acid. The acetic acid is distilled off, then fractional distillation of the residue under vacuum gives bromocresols from which 4-bromo-*m*-cresol is obtained by crystallisation from hexane. Addition of the bromocresol in glacial acetic acid slowly to a reaction mixture of HI and red phosphorus or (more smoothly) of HI and hypophosphorus acid, in glacial acetic acid, at reflux, removes the bromine. After an hour, the solution is distilled at atmospheric pressure until layers are formed. Then it is cooled and diluted with water. The cresol is extracted with ether, washed with water, NaHCO<sub>3</sub> solution and again with water, dried with a little CaCl<sub>2</sub> and distilled [Baltzly et al. *J Am Chem Soc* 77 2522 1955]. The *3,5-dinitrobenzoate* (prepared from 3,5-dinitrobenzoyl chloride in dry pyridine, and recrystallised from EtOH or aqueous Me<sub>2</sub>CO) has **m 165°**. [Beilstein 6 IV 2035.]

***p*-Cresol** [106-44-5] **M 108.1, m 34.8°, b 201.9°, n<sub>D</sub><sup>41</sup> 1.531, n<sub>D</sub><sup>46</sup> 1.529, pK<sup>25</sup> 10.27.** It can be separated from *m*-cresol by fractional crystallisation of its melt. Purify it by distillation, by precipitation from \*benzene solution with petroleum ether, and *via* its benzoate, as for phenol. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. It has also been crystallised from petroleum ether (b 40-60°) and by conversion to sodium *p*-cresoxyacetate which, after crystallisation from water is decomposed by heating with HCl in an autoclave [Savard *Ann Chim (Paris)* 11 287 1929]. The *3,5-dinitrobenzoate* (prepared from 3,5-dinitrobenzoyl chloride in dry pyridine, and recrystallised from EtOH or aqueous Me<sub>2</sub>CO) has **m 189°**. [Beilstein 6 II 2093.]

***o*-Cresolphthalein complexon (Metalphthalein)** [2411-89-4] **M 636.6, m 186°(dec),  $\lambda_{\max}$  575nm, pK<sub>1</sub> 2.2, pK<sub>2</sub> 2.9, pK<sub>3</sub> 7.0, pK<sub>4</sub> 7.8, pK<sub>5</sub> 11.4, pK<sub>6</sub> 12.0.** *o*-Cresolphthalein (a complexon precursor without the two bis-carboxymethylamino groups) is a contaminant and is one of the starting materials. It can be removed by dissolving the reagent in H<sub>2</sub>O and adding a 3-fold excess of sodium acetate and fractionally precipitating it by dropwise addition of HCl to the clear filtrate. Wash the precipitate with cold H<sub>2</sub>O and dry the *monohydrate* at 30° in a vacuum. The pure material gives a single spot on paper chromatography (eluting solvent EtOH/water/phenol, 6:3:1, and developing with NaOH). [Anderegg et al. *Helv Chim Acta* 37 113 1954.] It complexes with Ba, Ca, Cd, Mg and Sr. [Beilstein 18 III/IV 8141.]

***o*-Cresol Red** [1733-12-6] **M 382.4, m 290°(dec), pK<sup>25</sup> 1.26.** Crystallise it from glacial acetic acid. Dry it in air. Dissolve it in aqueous 5% NaHCO<sub>3</sub> solution and precipitate it from a hot solution by dropwise addition of aqueous HCl. Repeat the procedure till the UV maximum does not increase. [Beilstein 19 IV 1133.]

***o*-Cresotic acid (3-methylsalicylic acid)** [83-40-9] **M 152.2, m 163-164°, 165°, pK<sub>1</sub><sup>25</sup> 3.32.** Crystallise the acid from water. [Beilstein 10 H 220, 10 II 131, 10 III 505, 10 IV 601.]

***m*-Cresotic acid (4-methylsalicylic acid)** [50-85-1] **M 152.2, m 176°, 177°, (182-183°), pK<sub>1</sub><sup>25</sup> 3.15, pK<sub>2</sub><sup>25</sup> 13.35.** Crystallise the acid from water. It sublimes at 130°/11mm. [Beilstein 10 H 233, 10 II 137, 10 III 521, 10 IV 617.]

***p*-Cresotic acid (5-methylsalicylic acid)** [89-56-5] **M 152.2, m 151°, 152°, 151-154°, pK<sub>1</sub><sup>25</sup> 3.40, pK<sub>2</sub><sup>25</sup> 13.45.** Crystallise the acid from H<sub>2</sub>O. [Beilstein 10 H 227, 10 II 134, 10 III 516, 10 IV 610.]

**Crystal Violet Chloride {Gentian violet, *N*-4[bis[4-(dimethylaminophenyl)methylene]-2,5-cyclohexadien-1-ylidene]-*N*-methylmethaninium chloride}** [548-62-9] **M 408.0, pK 9.36.** Crystallise the dye from water (20ml/g), the crystals being separated from the chilled solution by centrifugation, then wash them with chilled EtOH (solubility is 1g in 10 ml of hot EtOH) and diethyl ether and dry under vacuum. It is soluble in CHCl<sub>3</sub> but insoluble in Et<sub>2</sub>O. The carbinol is precipitated from an aqueous solution of the dye-hydrochloride, using excess NaOH, then dissolve in HCl and recrystallise it from water as the chloride [UV and kinetics: Turgeon & La Mer *J Am Chem Soc* 74 5988 1952]. The *carbinol base* has **m 195°** (needles from EtOH). The *diphthalate* (blue and turns red in H<sub>2</sub>O) crystallises from H<sub>2</sub>O, **m 153-154°(dec at 185-187°)**[Chamberlain & Dull *J Am Chem Soc* 50 3089 1928]. [Beilstein 13 H 233, 13 IV 2284.]

**Cumene (isopropyl benzene)** [98-82-8] **M 120.2, b 69-70°/41mm, 152.4°/760mm, d<sub>4</sub><sup>20</sup> 0.864, n<sub>D</sub><sup>20</sup> 1.49146, n<sub>D</sub><sup>25</sup> 1.48892.** Usual purification is by washing it with several small portions of conc H<sub>2</sub>SO<sub>4</sub> (until the acid layer is no longer coloured), then with water, 10% aqueous Na<sub>2</sub>CO<sub>3</sub>, again with water, and drying with MgSO<sub>4</sub>, MgCO<sub>3</sub> or Na<sub>2</sub>SO<sub>4</sub>, followed by fractional distillation. It can then be dried with, and distilled from, Na, NaH or CaH<sub>2</sub>. Passage through columns of alumina or silica gel removes oxidation products. It has also been steam distilled from 3% NaOH, and azeotropically distilled with 2-ethoxyethanol (which is subsequently removed by washing out with water). [Beilstein 5 IV 985.]

**Cumene hydroperoxide** [80-15-9] **M 152.2, b 60°/0.2mm, d<sub>4</sub><sup>20</sup> 1.028, n<sub>D</sub><sup>24</sup> 1.5232.** Purify the hydroperoxide by adding 100ml of 70% material slowly and with agitation to 300ml of 25% NaOH in water, keeping the temperature below 30°. The resulting crystals of the sodium salt are filtered off, washed twice with 25 ml portions of \*benzene, then stirred with 100ml of \*benzene for 20 minutes. After filtering off the crystals and repeating the washing, they are suspended in 100ml of distilled water and the pH is adjusted to 7.5 by addition of 4M HCl. The free hydroperoxide is extracted into two 20ml portions of *n*-hexane, and the solvent is evaporated under vacuum at room temperature, the last traces being removed at 40-50°/1mm [Fordham & Williams *Canad J Res* 27B 943 1949]. Petroleum ether, **but not diethyl ether**, can be used instead of \*benzene, and powdered solid CO<sub>2</sub> can replace the 4M HCl. [Beilstein 6 IV 3221.] *The material is potentially EXPLOSIVE.*

**Cuminaldehyde (4-isopropylbenzaldehyde)** [122-03-2] **M 148.2, b 82-84°/3.5mm, 120°/23mm, 131-135°/35mm, 235-236°/760mm, d<sub>4</sub><sup>20</sup> 0.978, n<sub>D</sub><sup>20</sup> 1.5301.** A likely impurity is the benzoic acid. Check the IR for the presence of OH from CO<sub>2</sub>H, and the CO frequencies. If the acid is present, then dissolve the aldehyde in Et<sub>2</sub>O, wash it with 10% NaHCO<sub>3</sub> until effervescence ceases, then with brine, dry over CaCl<sub>2</sub>, evaporate and distil the residual oil, preferably under vacuum. It is almost insoluble in H<sub>2</sub>O, but soluble in EtOH and Et<sub>2</sub>O. The *thiosemicarbazone* has **m 147°** after recrystallisation from aqueous EtOH, MeOH or \*C<sub>6</sub>H<sub>6</sub>. [Crouse *J Am Chem Soc* 71 1263 1949, Bernstein et al. *J Am Chem Soc* 73 906 1951, Gensler & Berman *J Am Chem Soc* 80 4949 1958, Beilstein 7 H 318, 7 II 347, 7 III 1095, 7 IV 723.]

**9-Cyanoanthracene (anthracene-9-carbonitrile)** [1210-12-4] **M 203.2, m 134-137°, 173-177°.** Crystallise the nitrile from EtOH or toluene, and sublime it in a vacuum in the dark under N<sub>2</sub> [Ebied et al. *J Chem Soc, Faraday Trans 1* 76 2170 1980, Kikuchi et al. *J Phys Chem* 91 574 1987]. [Beilstein 9 I 304.]

**9-Cyanoanthracene photodimer** [33998-38-8] **M 406.4, dec to monomer above ~147°**. Purify the dimer by dissolving it in the minimum amount of  $\text{CHCl}_3$  followed by addition of EtOH at 5° [Ebied et al. *J Chem Soc, Faraday Trans 1* **75** 1111 1979, Ebied et al. *J Chem Soc, Faraday Trans 1* **76** 2170 1980].

*p*-Cyanobenzoic acid [619-65-8] **M 147.1, m 219°, 219-221° (dec), pK<sup>25</sup> 3.55**. Crystallise the acid from water and dry it in a vacuum desiccator over Sicapent. [Beilstein **9** IV 3324.]

**4-Cyanobenzoyl chloride** [6068-72-0] **M 165.6, m 68-70°, 69-70°, 73-74°, b 132°/8mm, 150-151°/25mm**. If the IR shows the presence of OH, then treat it with  $\text{SOCl}_2$  boil for 1 hour, evaporate and distil it in a vacuum. The distillate solidifies and can be recrystallised from petroleum ether. It is moisture sensitive and an **IRRITANT**. [Ashley et al. *J Chem Soc* 103 1942, Fison et al. *J Org Chem* **16** 648 1951, [Beilstein **9** III 4255, **14** IV 3327.]

*p*-Cyanophenol (*p*-hydroxybenzonitrile) [767-00-0] **M 119.1, m 113°, pK<sup>25</sup> 7.97**. Crystallise the phenol from petroleum ether, \*benzene or water and keep it under vacuum over  $\text{P}_2\text{O}_5$ . [Bernasconi & Paschelis *J Am Chem Soc* **108** 2969 1986.] [Beilstein **10** H 167, **10** IV 441.]

**Cyclohexylbenzene (phenylcyclohexane)** [827-52-1] **M 160.3, f 6.8°, b 237-239°, d<sub>4</sub><sup>20</sup> 0.950, n<sub>D</sub><sup>20</sup> 1.5258**. Purify it by fractional distillation, and by fractional freezing. [Beilstein **5** IV 1424.]

**Cyclopropyldiphenylcarbinol (cyclopropyldiphenylmethanol)** [5785-66-0] **M 224.3, m 86-87°**. Crystallise the carbinol from *n*-heptane or \* $\text{C}_6\text{H}_6$ /pentane (m 82-83°). It sublimes at 60°/0.001mm. The 2,4-dinitrobenzoate has m 140°. [Beilstein **6** III 3517, **6** IV 4888.]

*p*-Cymene (4-isopropyltoluene) [99-87-6] **M 134.2, b 177.1°/760mm, d<sub>4</sub><sup>20</sup> 0.8569, n<sub>D</sub><sup>20</sup> 1.4909, n<sub>D</sub><sup>25</sup> 1.4885**. Wash *p*-cymene with cold, conc  $\text{H}_2\text{SO}_4$  until there is no further colour change, then repeatedly with  $\text{H}_2\text{O}$ , 10% aqueous  $\text{Na}_2\text{CO}_3$  and  $\text{H}_2\text{O}$  again. Dry it over  $\text{Na}_2\text{SO}_4$ ,  $\text{CaCl}_2$  or  $\text{MgSO}_4$ , and distil it. Further purification steps include steam distillation from 3% NaOH, percolation through silica gel or activated alumina, and a preliminary reflux for several days over powdered sulfur. Store it over  $\text{CaH}_2$ . [Beilstein **5** IV 1060.]

**Deoxybenzoin** [451-40-1] **M 196.3, m 60°, b 177°/12mm, 320°/760mm**. Crystallise deoxybenzoin from EtOH and/or distil it in a vacuum. [Beilstein **7** II 368, **7** III 2098, **7** IV 1393.]

(±)-Desyl bromide ( $\alpha$ -bromo-desoxybenzoin,  $\omega$ -bromo- $\omega$ -phenyl acetophenone) [484-50-0] **M 275.2, m 57.1-57.5°**. Crystallise it from 95% EtOH. [Beilstein **7** H 436, **7** II 370, **7** III 2122.]

(±)-Desyl chloride ( $\alpha$ -chloro-desoxybenzoin,  $\omega$ -chloro- $\omega$ -phenyl acetophenone) [447-31-4] **M 230.7, m 62-64°, 66-67°, 67.5°, 68°**. For the purification of small quantities recrystallise it from petroleum ether (b 40-60°), but use MeOH or EtOH for larger quantities. For the latter solvent, dissolve 12.5g of chloride in 45ml of boiling EtOH (95%), filter and the filtrate yields colourless crystals (7.5g) on cooling. A further crop (0.9g) can be obtained by cooling in an ice-salt bath. It turns brown on exposure to sunlight but it is stable in sealed dark containers. The *R*(+)-enantiomer has m 75-76° (from petroleum ether) and  $[\alpha]_{546} +168.4^\circ$  (c 0.6,  $\text{Me}_2\text{CO}$ ) [Roger & Wood *J Chem Soc* 811 1954]. [Henley & Turner *J Chem Soc* 1182 1931, Ward *Org Synth Coll Vol II* 159 1943, Beilstein **7** H 436, **7** I 234, **7** II 369, **7** III 2106, **7** IV 1396.]

**Diacetyliodobenzene (iodobenzenediacetate)** [3240-34-4] **M 322.1, m 163-165°**. The purity of diacetyliodobenzene can be checked by treatment with  $\text{H}_2\text{SO}_4$  then KI, and the liberated  $\text{I}_2$  is estimated with standard thiosulfate. It has been recrystallised from 5M acetic acid and dried overnight in a vacuum desiccator over  $\text{CaCl}_2$ . The surface of the crystals may become slightly yellow but this does not affect its usefulness. [Sharefkin & Saltzman *Org Synth Coll Vol V* 600 1973, Beilstein **5** IV 693.]

**1,2-Diacetyl benzene** [704-00-7] **M 162.2, m 39-41°, 41-42°, b 110°/0.1mm, 148°/20mm**. Purify it by

distilling and by recrystallising from petroleum ether. The *bis-2,4-dinitrophenylhydrazone* has **m** 221° (dec). [Halford & Weissmann *J Org Chem* **17** 1646 1952, Riemschneider & Kassahn *Chem Ber* **92** 1705 1959, [Beilstein **7** III 3501, **7** IV 2155.]

**1,4-Diacetyl benzene** [1009-61-6] **M** 162.2, **m** 113-5-114.2°, **b** 128-130°/3mm. Crystallise it from EtOH (**m** 114°) or \*benzene and dry it in a vacuum over CaCl<sub>2</sub>. Also purify it by dissolving it in acetone, treating with Norit, evaporating and recrystallising from MeOH. The *dioxime* has **m** 248-259°. [Wagner et al. *J Am Chem Soc* **108** 7727 1986]. [Beilstein **7** H 686, **7** II 624, **7** III 3504, **7** IV 2156.]

**1,4-Diaminoanthraquinone** [128-95-0] **M** 238.3, **m** 268°. Purify the anthraquinone by thin-layer chromatography on silica gel using toluene/acetone (9:1) as eluent. The main band is scraped off and extracted with MeOH. The solvent is evaporated, and the quinone is dried in a drying pistol [Land et al. *J Chem Soc, Faraday Trans 1* **72** 2091 1976]. It crystallises from EtOH (**m** 269°) in dark violet crystals. Store it in sealed ampoules in the dark. [Beilstein **14** H 197, **14** II 113, **14** III 437, **14** IV 458.]

**1,5-Diaminoanthraquinone** [129-44-2] **M** 238.3, **m** 319°. Recrystallise it from aniline (**m** 313-314°), EtOH or acetic acid [Flom & Barbara *J Phys Chem* **89** 4481 1985]. [Beilstein **14** H 303, **14** I 467, **14** II 116, **14** III 466, **14** IV 479.]

**2,6-Diaminoanthraquinone** [131-14-6] **M** 238.3, **m** 310-320°. Crystallise it from pyridine or nitrobenzene (red needles). Column-chromatography on Al<sub>2</sub>O<sub>3</sub>/toluene is used to remove a fluorescent impurity, then it is recrystallised from EtOH. [Beilstein **14** H 215, **14** I 471, **14** II 120, **14** III 480, **14** IV 486.]

**3,3'-Diaminobenzidine tetrahydrochloride (2H<sub>2</sub>O)** [7411-49-6] **M** 396.1, **m** >300°(dec), **pK<sub>Est(1)</sub>** ~3.3, **pK<sub>Est(2)</sub>** ~4.7 (free base). Dissolve the salt in water and precipitate it by adding conc HCl, then drying it over solid NaOH. [Beilstein **13** IV 530.]

**3,4-Diaminobenzoic acid** [619-05-6] **M** 152.2, **m** 213°(dec), 228-229°, **pK<sub>1</sub><sup>25</sup>** 2.57 (4-NH<sub>2</sub>), **pK<sub>2</sub><sup>25</sup>** 3.39 (3-NH<sub>2</sub>), **pK<sub>Est(3)</sub>** ~5.1 (CO<sub>2</sub>H). Crystallise it from H<sub>2</sub>O or toluene. [Beilstein **15** IV 1503.]

**3,5-Diaminobenzoic acid** [535-87-5] **M** 152.2, **m** 235-240°(dec), **pK<sup>25</sup>** 5.13 (CO<sub>2</sub>H), **pK<sup>25</sup>** 7.12 (in 80% aqueous 2-MeOCH<sub>2</sub>CH<sub>2</sub>OH). Crystallise the acid from water. The *dihydrochloride* has **m** 226-228°(dec). [Beilstein **14** H 453, **14** III 1179, **14** IV 1304.]

**3,4-Diaminobenzophenone** [39070-63-8] **M** 212.3, **m** 116-117°, **pK<sub>Est(1)</sub>** ~<0, **pK<sub>Est(2)</sub>** ~2.5. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and sublime it *in vacuo* [Ayyanger et al. *Org Prep Proced Int* **23** 627 1991].

**4,4'-Diaminobenzophenone** [611-98-3] **M** 212.3, **m** 242-244°, 243-245°, 246.5-247.5° (after sublimation at 0.0006mm), **pK<sub>1</sub><sup>25</sup>** 1.37, **pK<sub>2</sub><sup>25</sup>** 2.92. Purify the phenone by recrystallisation from EtOH and by sublimation in high vacuum. The *dihydrochloride* has **m** 260°(dec) (from EtOH), and the *thiosemicarbazone* has **m** 207-207.5°(dec) (from aqueous EtOH). [Kuhn et al. *Chem Ber* **75** 711 1942, Beilstein **14** IV 255.]

**1,2-Diamino-4,5-dichlorobenzene** [5348-42-5] **M** 177.0, **m** 162°, 162-163°, 163°, **pK<sub>Est(1)</sub>** ~1.0, **pK<sub>Est(2)</sub>** ~2.9. Reflux the amine with activated charcoal in CH<sub>2</sub>Cl<sub>2</sub>, followed by recrystallisation from Et<sub>2</sub>O/petroleum ether or petroleum ether [Koolar & Kochi *J Org Chem* **52** 4545 1987]. Alternatively, recrystallise the diamine from hexane, \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or H<sub>2</sub>O (Na<sub>2</sub>SO<sub>4</sub>), and sublime it at 150°/15mm. [Beilstein **13** IV 72.]

**4,4'-Diamino-3,3'-dinitrobiphenyl** [6271-79-0] **M** 274.2, **m** 275°, **pK<sub>Est</sub>** ~-0.2. Crystallise the biphenyl from aqueous EtOH or pyridine/EtOH. [Beilstein **13** H 236, **13** I 68, **13** II 108, **13** III 415, **13** IV 387.]

**4,4'-Diaminodiphenylamine** [537-65-5] **M** 199.3, **m** 158°, **pK<sub>Est</sub>** ~5.0. Crystallise the amine from water (its solubility at 25° is 0.42%). [Beilstein **13** H 110, **13** I 36, **13** II 55, **13** III 256, **13** IV 203.]

**4,4'-Diaminodiphenyl ether (4,4'-oxydianiline, DADPE)** [101-80-4] **M** 200.2, **m** 185°, 186-187°, 188-192°,

**pK<sup>25</sup> 5.12 (50% EtOH).** This base has been prepared by several syntheses including, condensation of *p*-nitrophenol with *p*-chloroaniline, nitration of diphenyl ether to 4,4'-dinitrodiphenyl ether and reduction with hydrazine and Raney Ni [Shamis & Dashevski *J Org Chem, USSR Engl Transl* **3** 1005 1967], a Friedel-Crafts reaction of diphenyl ether (AcCl, AlCl<sub>3</sub>) to 4,4'-diacetyldiphenyl ether (m 99-100°), conversion to the dioxime (m 180-183°), and a bis-Beckman rearrangement (AcOH, Ac<sub>2</sub>O, HCl gas) gives 4,4'-bis-acetamidodiphenyl ether (90%). This amide (46g) in H<sub>2</sub>O (600ml) and concentrated HCl (60ml) is refluxed for 2 hours until it all dissolved, filtered (charcoal), the filtrate is basified with Na<sub>2</sub>CO<sub>3</sub> solution, the precipitate is filtered off, washed with H<sub>2</sub>O, and dried at 100-110° to give the *free base* (32g). It is purified by zone melting and/or by sublimation at high vacuum. In the UV it has λ<sub>max</sub> at 248 and 300nm in EtOH. [Franovskii et al. *J Org Chem, USSR Engl Transl* **6** 2315 1970.] The *benzenesulfonate salt* crystallises from MeOH with **m 286-288°(dec)**. [Beilstein **13** H 441, **13** I 148, **13** III 106, **13** IV 1038.]

**4,4'-Diaminodiphenylmethane** [101-77-9] **M 198.3, m 91.6-92°, 92-93°, b 232°/9mm, pK<sub>Est</sub> ~4.9.** Crystallise the amine from water, 95% EtOH or \*benzene. [Beilstein **13** IV 390.]

**4,4'-Diaminodiphenyl sulfide (4,4'-thioaniline)** [139-65-1] **M 216.3, m 108-109°, pK<sup>20</sup> 2.28 (1:1 EtOH/H<sub>2</sub>O).** It separates as needles from EtOH and is a **possible mutagen**. The free base is used for the detection of NO<sub>3</sub> ions. The *diacetate* crystallises from aqueous AcOH with **m 182°** and the *sulfoxide*, [119-59-5] **m 184°**, forms prisms from EtOH or H<sub>2</sub>O. [Fuson & Melamed *J Org Chem* **13** 690 11948, *Beilstein* **13** III 1246, **13**, IV 1246.]

**2,7-Diaminofluorene** [524-64-4] **M 196.3, m 165°, pK<sub>Est</sub> ~4.6.** Recrystallise it from H<sub>2</sub>O. [Beilstein **13** IV 449.]

**1,5-Diaminonaphthalene** [2243-62-1] **M 158.2, m 190°, pK<sup>25</sup> 4.12.** Recrystallise the amino-naphthalene from boiling H<sub>2</sub>O, but this is wasteful due to poor solubility. Boil it in chlorobenzene (charcoal), filter hot and cool the filtrate (preferably under N<sub>2</sub>). This gives colourless crystals. Dry it in a vacuum till free from chlorobenzene (odour), and store it in sealed ampoules under N<sub>2</sub> away from light. [Beilstein **13** IV 340.]

**1,8-Diaminonaphthalene** [479-27-6] **M 158.2, m 66.5°, b 205°/12mm, pK<sup>25</sup> 4.44.** Crystallise 1,8-diaminonaphthalene from water or aqueous EtOH, and sublime it in a vacuum. The *N,N'*-dimethyl derivative [20734-56-9] has **m 103-104°** and **pK<sup>25</sup> 5.61**, the *N,N,N'*-trimethyl- derivative [20734-57-0] has **m 29-30°** and **pK<sup>25</sup> 6.43**. [Hodgson et al. *J Chem Soc* 202 1945, *Beilstein* **13** IV 344.]

**2,3-Diaminonaphthalene** [771-97-1] **M 158.2, m 199°, pK<sup>21</sup> 3.54 (in 50% aqueous EtOH).** Crystallise the diamine from water, or dissolve it in 0.1M HCl, by heating to 50°. After cooling, the solution is extracted with decalin to remove fluorescent impurities and centrifuged. [Beilstein **13** IV 346.]

**2,5-Di-tert-amylhydroquinone** [79-74-3] **M 250.4, m 185.8-186.5°.** Crystallise the hydroquinone under N<sub>2</sub> from boiling AcOH (7ml/g) plus boiling water (2.5ml/g). [Stolow & Bonaventura *J Am Chem Soc* **85** 3636 1963]. Store it in sealed ampoules under N<sub>2</sub> away from light. [Beilstein **6** H 952, **6** III 4748.]

**Di-n-amyl phthalate (dipentyl phthalate)** [131-18-0] **M 306.4, b 204-206°/11mm, d<sub>4</sub><sup>25</sup> 1.023, n<sub>D</sub><sup>20</sup> 1.489.** Wash the ester with aqueous Na<sub>2</sub>CO<sub>3</sub>, then distilled water. Dry it with CaCl<sub>2</sub> and distil it in a vacuum. Store it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. [Beilstein **9** IV 3178.]

**Diazoaminobenzene (1,3-diphenyltriazene)** [136-36-6] **M 197.2, m 99°, 100°.** Crystallise the triazene from \* petroleum ether (b 60-80°) (**m 94-96°**), 60% MeOH/water or 50% aqueous EtOH (charcoal) containing a small amount of KOH. Its solubility in petroleum ether (b 60-80°) is ~6%. Also purify it by chromatography on alumina/toluene and elute with toluene/petroleum ether. Store the pale yellow needles in the dark. [Hartman & Dickey *Org Synth Coll Vol II* 163 1943, *Beilstein* **16** H 687, **16** I 404, **16** II 351, **16** III 643, **16** IV 904.]

**Dibenzalacetone [1,5-diphenyl-1,4-dien-3-one, 1,5-(bisphenyl)-penta-1E,4E-diene-3-one]** [538-58-9] **M 234.3, m 107-113°, 112°, 120-122°.** Purify the ketone by flash chromatography (150 mesh Al<sub>2</sub>O<sub>3</sub> deactivated with 6% v/w H<sub>2</sub>O) using petroleum ether 40-60°/EtOAc (4/1, v/v), and the yellow solid is recrystallised by

layering a concentrated  $\text{CH}_2\text{Cl}_2$  solution with hexane (i.e.  $\text{CH}_2\text{Cl}_2/\text{hexane}$ , 1:4, v/v). Also recrystallise the ketone from hot ethyl acetate (2.5ml/g) or EtOH. The IR ( $\text{CH}_2\text{Cl}_2$ ) has  $\nu_{\text{max}}$  at 1657m, 1651m (C=O), 1627vs (C=C), 1591w (C=C aromatic), 1574w (C=C aromatic), 983m (CH trans)  $\text{cm}^{-1}$ ; and UV has  $\lambda_{\text{max}}$  (THF) at 233 ( $\pi-\pi^*$ ) and 321 (n- $\pi^*$ ) nm; the  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 7.65-7.15 (m, 10 H, ArH), 4.4 (dd,  $J = 5.8$ , 10.0 Hz, 1 H, C3a-H), 3.45-3.30 (m, 1 H, C6-H), 3.15-3.00 (m, 1 H, C6-H), 1.90-1.55 (m, 3 H, C4-H, C5-H<sub>2</sub>), 0.95-0.75 (m, 1 H, C4-H), 0.40 (s, 3 H, B-CH<sub>3</sub>). [Beilstein 7 IV 1747.]

**Dibenz[*a,h*]anthracene (1,2:5,6-dibenzanthracene)** [53-70-3] M 278.4, m 266-267°. The yellow-green colour (due to other pentacyclic impurities) is removed from it by crystallising from  $^b$ benzene or by selective oxidation with lead tetraacetate in acetic acid [Moriconi et al. *J Am Chem Soc* 82 3441 1960]. [Beilstein 5 IV 2722.]

**trans-1,2-Dibenzoyl ethylene (trans-1,4-diphenyl-2-butane-1,4-dione)** [959-28-4] M 236.3, m 109-112°, 111°. It crystallises from MeOH or EtOH as yellow needles [Koller et al. *Helv Chim Acta* 29 512 1946]. The dioxime has m 210-211°(dec) from AcOH. [IR: Kuhn et al. *J Am Chem Soc* 72 5058 1950, Yates *J Am Chem Soc* 74 5375 1952, Erickson et al. *J Am Chem Soc* 73 5301 1951, Beilstein 7 IV 2578.]

**Dibenzoylmethane (1,3-diphenyl-1,3-propanedione)** [120-46-7] M 224.3, m 80°. Crystallise dibenzoylmethane from petroleum ether or MeOH. [Beilstein 7 IV 2512.]

**2,3,6,7-Dibenzphenanthrene (pentaphene)** [222-93-5] M 276.3, m 255-256°, 257°. Purify pentaphene through  $\text{Al}_2\text{O}_3$  with  $^*\text{C}_6\text{H}_6$  or xylene as eluant. It crystallises from xylene as yellow plates and sublimes in high vacuum. The dipicrate forms orange needles m 184° from  $^*\text{C}_6\text{H}_6$ . (Clar & John *Ber* 64 986 1931, Clar & Stewart *J Chem Soc* 3215 1951, *Tei Lett* 3023 1965, French & Zander *Ber* 99 396 1966.)

**Dibenzyl amine** [103-49-1] M 197.3, m -26°, b 113-114°/0.1mm, 174-175°/6mm, 270°/250mm, 300° (partial dec),  $d_4^{20}$  1.027,  $n_D^{20}$  1.576,  $\text{pK}^{25}$  8.52. Purify the amine by distillation in a vacuum. It causes burns to the skin. The dihydrochloride has m 265-266° (from MeOH/HCl), and the tetraphenyl boronate has m 129-133°. [Bradley & Maisey *J Chem Soc* 247 1954, Hall *J Phys Chem* 60 63 1956, Donetti & Bellora *J Org Chem* 37 3352 1972, Beilstein 12 IV 2179.]

**Dibenzyl disulfide** [150-60-7] M 246.4, m 71-72°, 74-75°, b 142-148°/0.05-0.1mm, 210-216°/18mm. Crystallise the disulfide from EtOH (m 77°), petroleum ether or  $\text{CS}_2$  (m 72°) or distil it. The  $\text{AgNO}_3$  complex has m 103°. [Beilstein 6 H 465, 6 I 229, 6 II 437, 6 III 1635, 6 IV 2760.]

**Dibenzylethylenediamine (benzathine, DBED)** [140-28-3] M 240.4, m 26°, b 195°/4mm,  $d_4^{20}$  1.02,  $n_D^{20}$  1.563,  $\text{pK}_{\text{Est}(1)} \sim 5.9$ ,  $\text{pK}_{\text{Est}(2)} \sim 8.9$ . Dissolve DBED in acid, extract with toluene, basify, extract it with  $\text{Et}_2\text{O}$ , dry over solid KOH, evaporate and fractionate it *in vacuo*. The diacetate crystallises from  $\text{H}_2\text{O}$  by addition of EtOH and has m 111° (solubility in  $\text{H}_2\text{O}$  is ~25%). The dihydrochloride has m 306-308° (from  $\text{H}_2\text{O}$ ) and the dipicrate has m 212°(dec) (from  $\text{H}_2\text{O}$ ). [Frost et al., *J Am Chem Soc* 71 3842 1949, Beilstein 12 H 1067, 12 III 2304.]

**Dibenzyl ketone (1,3-diphenyl-2-propanone)** [102-04-5] M 210.3, m 34.0°. Fractionally crystallise it from its melt, then crystallise it from petroleum ether. Store it in the dark. [Beilstein 7 IV 1420.]

**Dibenzyl malonate** [15014-25-2] M 284.3, b 188-190°/0.2mm, 193-196°/1mm,  $d_4^{20}$  1.158,  $n_D^{20}$  1.5452. Dissolve the ester in toluene, wash it with aqueous  $\text{NaHCO}_3$ ,  $\text{H}_2\text{O}$ , dry over  $\text{MgSO}_4$ , filter, evaporate and distil it at high vacuum. [Ginsburg & Pappo *J Am Chem Soc* 75 1094 1953, Baker et al. *J Org Chem* 17 77 1952, Beilstein 6 IV 2270.]

**Dibenzyl sulfide** [538-74-9] M 214.3, m 48.5°, 50°. Crystallise the sulfide from EtOH/water (10:1), or repeatedly from  $\text{Et}_2\text{O}$ . It has also been purified by chromatography on  $\text{Al}_2\text{O}_3$  (pentane as eluent), then recrystallised from EtOH [Kice & Bowers *J Am Chem Soc* 84 2390 1962]. Dry it in a vacuum at 30° over  $\text{P}_2\text{O}_5$ , fused under nitrogen and re-dried. [Beilstein 6 IV 2649.]

**2,4-Dibromoaniline** [615-57-6] **M 250.9, m 79-80°, pK<sup>25</sup> 1.87.** Crystallise the aniline from aqueous EtOH. The *picrate* has **m 124°**. [Beilstein 12 H 655, 12 I 326, 12 II 356, 12 III 1471, 12 IV 1532.]

**9,10-Dibromoanthracene** [523-27-3] **M 336.0, m 222-224°, 226°.** Recrystallise it from toluene, xylene or CCl<sub>4</sub> (yellow needles), and sublime it in a vacuum [Johnston et al. *J Am Chem Soc* 109 1291 1987]. [Heilbron & Heaton *Org Synth Coll Vol I* 207 1941, *Beilstein* 5 H 665.]

***p*-Dibromobenzene** [106-37-6] **M 235.9, m 87.8°.** Steam distil the dibromobenzene, then crystallise it from EtOH or MeOH and dry it in the dark under vacuum. Purify it by zone melting. [Beilstein 5 IV 683.]

**2,5-Dibromobenzoic acid** [610-71-9] **M 279.9, m 157°, pK<sub>Est</sub> ~1.5.** Crystallise the acid from water or EtOH. [Beilstein 9 H 358, 9 I 147, 9 II 237, 9 III 1428, 9 IV 1027.]

**4,4'-Dibromobiphenyl** [92-86-4] **M 312.0, m 164°, b 355-360°/760mm.** Crystallise it from MeOH. [Beilstein 5 IV 1820.]

**(±)-2,2'-Dibromo-1,1'-binaphthyl** [74866-28-7, (=) 76284-65-6] **M 412.1, m 178-183°, 180.5-181°.** Purify the binaphthyl by chromatography through a silica gel column (70-230mesh) using hexane as eluent. It gives pale yellow crystals from EtOH with **m 187.3-187.9°**. The *R*-(+)-*enantiomer* [86688-08-6] crystallises from hexane with **m 157-157.5°** and **[α]<sub>D</sub><sup>25</sup> +32.9°** (c 1, pyridine) [Brown et al. *J Org Chem* 50 4345 1985]. [Okamoto et al. *J Am Chem Soc* 103 6971 1981, *Beilstein* 5 III 2465.]

**α,α-Dibromodeoxybenzoin** [15023-99-1] **M 354.0, m 111.8-112.7°.** Crystallise α,α-dibromo deoxybenzoin from acetic acid, 50% aqueous EtOH (prisms, **m 109-112°**) or Et<sub>2</sub>O (**m 112°**). [Curtius & Lang *J Prakt Chem* 44 547 1891, *Beilstein* 7 H 436, 7 III 2114.]

**2,5-Dibromonitrobenzene** [3460-18-2] **M 280.9, m 84°, 85-86°.** It crystallises from Me<sub>2</sub>CO or EtOH. [Beilstein 5 H 250, 5 II 190, 5 III 621, 5 IV 732.]

**2,6-Dibromo-4-nitrophenol** [99-28-5] **M 280.9, m 143-144°, pK<sup>25</sup> 3.39.** Crystallise the phenol from aqueous EtOH, 50% aqueous AcOH (**m 144-145°**, dec varies with rate of heating). Dry it in an oven at 40-60° or *in vacuo* over NaOH. [Hartman & Dickey *Org Synth Coll Vol II* 173 1943, *Beilstein* 6 H 246, 6 I 123, 6 II 234, 6 III 849, 6 IV 1366.]

**2,4-Dibromophenol** [615-58-7] **M 251.9, m 37°, 41-42°, b 154°/10mm, 239°/760mm, pK<sup>25</sup> 7.79.** Crystallise the phenol from CHCl<sub>3</sub> at -40°, or distil it in a vacuum. [Beilstein 6 H 202, 6 I 106, 6 II 188, 6 III 753, 6 IV 1061.]

**2,6-Dibromophenol** [608-33-3] **M 251.9, m 56-57°, b 138°/10mm, 255-256°/740mm, pK<sup>25</sup> 6.67.** Distil the phenol under vacuum (at 10mm), then crystallise it from cold CHCl<sub>3</sub> or from EtOH water. [Beilstein 6 H 202, 6 I 106, 6 II 188, 6 III 755, 6 IV 1064.]

**α,α'-Dibromo-*o*-xylene** [91-13-4] **M 264.0, m 95°, 95-96°, 98-99°, b 129-130°/4.5mm.** Crystallise it from CHCl<sub>3</sub> or petroleum ether, and/or distil it under vacuum. [Wenner *Org Chem* 17 527 1952, *Beilstein* 5 H 366, 5 I 180, 5 II 285, 5 III 819, 5 IV 929.]

**α,α'-Dibromo-*m*-xylene** [626-15-3] **M 264.0, m 77°, 78-79°, b 156-160°/12mm.** Crystallise it from acetone or \*benzene, and fractionally distil it under vacuum. [Wenner *Org Chem* 17 527 1952, *Beilstein* 5 H 374, 5 I 184, 5 II 294, 5 III 839, 5 IV 946.]

**α,α'-Dibromo-*p*-xylene** [623-24-5] **M 264.0, m 143.4°, 142-144°, 145-147°, b 155-158°/12-15mm, 245°/760mm.** Distil it under a vacuum and recrystallise it from EtOH, \*benzene or chloroform. [Wenner *Org Chem* 17 527 1952, *Beilstein* 5 H 385, 5 I 187, 5 II 301, 5 III 859, 5 IV 970.]

**$\alpha$ -Dibutylamino- $\alpha$ -(*p*-methoxyphenyl)acetamide (Ambucetamide)** [519-88-0] **M 292.4, m 134°**. Crystallise ambucetamide from EtOH containing 10% diethyl ether. [Janssen *J Am Chem Soc* **76** 6192 1954, *Beilstein* **14** IV 2101.]

**2,5-Di-*tert*-butyl aniline** [21860-03-7] **M 205.4, m 103-104°, 103-106°, pK<sup>25</sup> 3.34 (50% aqueous MeOH), 3.58 (90% aqueous MeOH)**. The aniline recrystallises from EtOH in fine needles after steam distillation. It has a pK<sup>25</sup> of 3.58 (50% aqueous EtOH). The *tosylate* has **m 164°** (from AcOH). [Bell & Wilson *J Chem Soc* 2340 1956, Carpenter et al. *J Org Chem* **16** 586 1951, Bartlett et al. *J Am Chem Soc* **76** 2349 1954, *Beilstein* **12** IV 2891.]

***p*-Di-*tert*-butylbenzene** [1012-72-2] **M 190.3, m 80°, 236°/760mm**. Crystallise it from Et<sub>2</sub>O or EtOH and dry it under vacuum over P<sub>2</sub>O<sub>5</sub> at 55°. [Tanner et al. *J Org Chem* **52** 2142 1987, *Beilstein* **5** II 344.]

**2,6-Di-*tert*-butyl-1,4-benzoquinone** [719-22-2] **M 220.3, m 66-67°**. It can be recrystallised from MeOH and sublimes in a vacuum. [*Beilstein* **7** IV 2116.]

**3,5-Di-*tert*-butyl-*o*-benzoquinone** [3383-21-9] **M 220.3, m 112-114°, 113-114°**. It can be recrystallised from MeOH or petroleum ether, and forms fine red plates or rhombs. [Flaig et al. *Justus Liebigs Ann Chem* **597** 196 1955, IR: Ley & Müller *Chem Ber* **89** 1402 1956, *Beilstein* **7** IV 2113.]

**3,5-Di-*tert*-butyl catechol** [1020-31-1] **M 222.3, m 99°, 99-100°, pK<sub>Est(1)</sub>~11.0, pK<sub>Est(2)</sub>~13.1**. Recrystallise the catechol from petroleum ether. [Ley & Müller *Chem Ber* **89** 1402 1956, UV Flaig et al. *Z Naturforschung* **10b** 668 1955.] Also purify it by crystallising three times from pentane [Funabiki et al. *J Am Chem Soc* **108** 2921 1986].

**2,6-Di-*tert*-butyl-*p*-cresol (2,6-di-*tert*-butyl-4-methylphenol, butylatedhydroxytoluene, BHT or DBPC)** [128-37-0] **M 230.4, m 71.5°, pK<sup>25</sup> 12.23**. Dissolve BHT in *n*-hexane at room temperature, then cool with rapid stirring, to -60°. The precipitate is separated, redissolved in hexane, and the process is repeated until the mother liquor is no longer coloured. The final product is stored under N<sub>2</sub> at 0° [Blanchard *J Am Chem Soc* **82** 2014 1960]. It has also been recrystallised from EtOH, MeOH, \*benzene, *n*-hexane, methylcyclohexane or petroleum ether (b 60-80°), and is dried in a vacuum. [*Beilstein* **6** IV 3511.]

**2,6-Di-*tert*-butyl-4-dimethylaminomethylphenol** [88-27-7] **M 263.4, m 93-94°, b 172°/30mm, pK<sub>Est</sub> ~12.0**. Crystallise it from *n*-hexane. [*Beilstein* **13** IV 2014.]

**Di-*tert*-butyldiperphthalate** [2155-71-7] **M 310.3, m (48°) 57-57.5°, decomposes at 108°**. Crystallise the perphthalate from Et<sub>2</sub>O or petroleum ether and dry it over H<sub>2</sub>SO<sub>4</sub>. The IR has  $\nu_{\max}$  at 1772cm<sup>-1</sup> in CCl<sub>4</sub>. [Milas & Surgemor *J Am Chem Soc* 68 642 1946, Milas & Kelin *J Org Chem* **36** 2900 1971, *Beilstein* **9** III 4190, **9** IV 3260.] **CARE**, potentially **EXPLOSIVE**.

**2,6-Di-*tert*-butyl-4-ethylphenol** [4130-42-1] **M 234.4, m 42-44°, pK<sub>Est</sub> ~12.3**. Crystallise the phenol from aqueous EtOH or *n*-hexane. [*Beilstein* **6** IV 3529.]

**2,5-Di-*tert*-butylhydroquinone** [88-58-4] **M 222.3, m 222-223°**. Crystallise the hydroquinone from \*C<sub>6</sub>H<sub>6</sub> or AcOH. [*Beilstein* **6** III 4741.]

**2,6-Di-*tert*-butyl-4-isopropylphenol** [5427-03-2] **M 248.4, m 39-41°, 38-42°, b 105-106°/0.3mm, pK<sub>Est</sub> ~12.3**. Crystallise the phenol from *n*-hexane or aqueous EtOH. It is used for making the respective *phenoxy* radical. [Cook & Norcross *J Am Chem Soc* **78** 3797 1956, *Beilstein* **6** III 3534.]

**2,6-Di-*tert*-butylphenol** [128-39-2] **M 206.3, m 37-38°, pK<sup>25</sup> 11.70**. Crystallise the phenol from aqueous EtOH or *n*-hexane. [*Beilstein* **6** III 2061.]

**Dibutyl phthalate (DBP, butyl phthalate)** [84-74-2] **M 278.4, f -35°, b 44°/2.5x10<sup>-4</sup>mm, 182°/5mm,**

**206°/20mm, 340°/760mm,  $d_4^{20}$  1.4929,  $d_5^5$  1.0426,  $n_D^{25}$  1.490.** Wash DBP with H<sub>2</sub>O (to free it from alcohol), then dilute NaOH (to remove any butyl hydrogen phthalate or acid), aqueous NaHCO<sub>3</sub> (charcoal), then distilled water. Dry it (CaCl<sub>2</sub>), distil it at 10torr or less, and store it in a desiccator over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* 9 II 586, 9 III 4102, 9 IV 3175.]

**Dichloramine-T (*N,N*-dichloro-*p*-toluenesulfonamide)** [473-34-7] **M 240.1, m 83°.** Crystallise it from petroleum ether (b 60-80°) or CHCl<sub>3</sub>/petroleum ether. Dry it in air <55° and store in the dark. It is soluble in CHCl<sub>3</sub> (~1:1), \*C<sub>6</sub>H<sub>6</sub> (~1:1) and CCl<sub>4</sub> (~1:2.5). It is a germicide and antibacterial as it readily liberates chlorine, and it should not smell strongly of chlorine; otherwise it should be purified. *Alternatively*, dissolve ~15g of reagent in 75ml of hot acetic acid and precipitate it by addition of 37ml of N/10 bleaching powder solution (NaOCl), filter off, wash it with a dilute solution of the latter, dry it as above (m 78-84°) and recrystallise it. Excessive drying, "Sharp drying", will decompose it. [Orton & Bradfield *J Chem Soc* 993 1927, Krauss & Crede *J Am Chem Soc* 39 2720 1917, Soper *J Chem Soc* 1899 1924; see also chloramine-T (monochloramine T Na salt) in "Compounds of As, B, P, S, Si." in this Chapter.] [*Beilstein* 11 H 107, 11 I 27, 11 II 63, 11 III 301.]

**2,4-Dichloroaniline** [554-00-7] **M 162.0, m 63°, pK<sup>25</sup> 2.02.** Crystallise the aniline from EtOH/water. It also crystallises from EtOH and is dried *in vacuo* for 6 hours at 40° [Moore et al. *J Am Chem Soc* 108 2257 1986, Edidin et al. *J Am Chem Soc* 109 3945 1987]. [*Beilstein* 12 IV 1241.]

**3,4-Dichloroaniline** [95-76-1] **M 162.0, m 71.5°, pK<sup>25</sup> 2.97.** Crystallise the aniline from MeOH. [*Beilstein* 12 IV 1257.]

**9,10-Dichloroanthracene** [605-48-1] **M 247.1, m 214-215°.** Purify it by crystallising it from MeOH, EtOH, \*C<sub>6</sub>H<sub>6</sub> or Me<sub>2</sub>CO (m 210-211°) followed by subliming *in vacuo*. [Masnori & Kochi *J Am Chem Soc* 107 7880 1985, *Beilstein* 5 H 664, 5 I 324, 5 II 575, 5 III 2134, 5 IV 2293.]

**2,4-Dichlorobenzaldehyde** [874-42-0] **M 175.0, m 72°.** Crystallise the aldehyde from EtOH or ligroin. [*Beilstein* 7 IV 575.]

**2,6-Dichlorobenzaldehyde** [83-38-5] **M 175.0, m 70.5-71.5°.** Crystallise the aldehyde from EtOH/H<sub>2</sub>O or petroleum ether (b 30-60°). [*Beilstein* 7 IV 576.]

***o*-Dichlorobenzene (ODCB)** [95-50-1] **M 147.0, b 81-82°/31-32mm, 180.5°/760mm,  $d_4^{20}$  1.306,  $n_D^{20}$  1.551,  $n_D^{25}$  1.549.** Contaminants may include the *p*-isomer and trichlorobenzene [Suslick et al. *J Am Chem Soc* 106 4522 1984]. It should be shaken with conc or fuming H<sub>2</sub>SO<sub>4</sub>, washed with water, dried with CaCl<sub>2</sub>, and distilled from CaH<sub>2</sub> or sodium in a glass-packed column. Low conductivity material (*ca* 10<sup>-10</sup> mhos) has been obtained by refluxing with P<sub>2</sub>O<sub>5</sub>, fractionally distilling and passing it through a column packed with silica gel or activated alumina: it is stored in a dry-box under N<sub>2</sub> or with activated alumina. [*Beilstein* 5 IV 654.]

***m*-Dichlorobenzene** [541-73-1] **M 147.0, b 173.0°/atm,  $d_4^{20}$  1.289,  $n_D^{20}$  1.54586,  $n_D^{25}$  1.54337.** Wash it with aqueous 10% NaOH, then with water until neutral, dry and distil it. Conductivity material (*ca* 10<sup>-10</sup> mhos) has been prepared by refluxing over P<sub>2</sub>O<sub>5</sub> for 8 hours, then fractionally distilling, and storing with activated alumina. *m*-Dichlorobenzene dissolves rubber stoppers. [*Beilstein* 5 IV 657.]

***p*-Dichlorobenzene** [106-46-7] **M 147.0, m 53.0°, b 174.1°,  $d_4^{20}$  1.241,  $n_D^{60}$  1.52849.** *o*-Dichlorobenzene is a common impurity. The *p*-isomer has been purified by steam distillation, crystallisation from EtOH or boiling MeOH, air-dried and dried in the dark under vacuum. It has also been purified by zone refining. [*Beilstein* 5 IV 658.]

**2,2'-Dichlorobenzidine** [84-68-4] **M 253.1, m 165°, pK<sub>Est(1)</sub> ~3.0, pK<sub>Est(2)</sub> ~4.0.** Crystallise the benzidine from EtOH or H<sub>2</sub>O. [*Beilstein* 13 H 234, 13 I 66, 13 II 106, 13 III 477, 13 IV 384.]

**3,3'-Dichlorobenzidine** [91-94-1] **M 253.1, m 132-133°, pK<sub>Est(1)</sub> ~4.8, pK<sub>Est(2)</sub> ~5.7.** Crystallise the benzidine from EtOH, petroleum ether (m 133°) or \*benzene. [*Beilstein* 13 H 234, 13 I 67, 13 II 106, 13 III 477,

## 13 IV 384.] CARCINOGEN.

**2,3-Dichlorobenzoic acid** [50-45-3] M 191.0, m 168.3°, 169-170°, pK<sup>25</sup> 2.67. Aromatic acid impurities (to <0.05%) can be removed *via* the (±)-α-methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* **12** 120 2008.] Crystallise the acid from H<sub>2</sub>O, aqueous EtOH, or 30% aqueous AcOH and several times from \*C<sub>6</sub>H<sub>6</sub>, then dry it *in vacuo* at 40° overnight. The *methyl ester* has m 35-39°. [Hope & Riley *J Chem Soc* **123** 2478 1923, Lock *Monatsh Chem* **90** 687 1959, Shorter & Mather *J Chem Soc* 4744 1961, *Beilstein* **9** II 228, **9** IV 998.]

**2,4-Dichlorobenzoic acid** [50-84-0] M 191.0, m 163-164°, pK<sup>25</sup> 2.68. Crystallise the acid from aqueous EtOH (charcoal), then \*benzene (charcoal). It can also be recrystallised from water. [*Beilstein* **9** IV 998.] It can be freed from isomeric acids (to <0.05%) *via* the (±)-α-methylbenzylamine salt as follows: dissolve the dichloro-acid (10g, 50.2mmol) in isopropanol (200ml), heat to 60° and add the (±)-benzylamine (5.49g, 45.3mmol), then stir it at 60° for 1 hour. Cool the mixture to room temperature, filter the slurry, wash it with isopropanol (25ml) and dry it *in vacuo* at 40° overnight to give 79% of the *salt* with m 185.2°. Dissolve the salt (5g) in H<sub>2</sub>O (50ml) and MeOH (20ml), then heat to 60° and add concentrated HCl to pH <2.0. Cool the solution to room temperature add H<sub>2</sub>O (12ml), filter it, wash it with H<sub>2</sub>O (30ml) and dry it *in vacuo* at 40° overnight to give 94% of the *acid* with m 162.0°. [Ley & Yates *Organic Process Research & Development* **12** 120 2008.]

**2,5-Dichlorobenzoic acid** [50-79-3] M 191.0, m 154°, b 301°/760mm, pK<sup>25</sup> 2.47. Crystallise the acid from water. [*Beilstein* **9** IV 1005.] Aromatic acid impurities (to <0.05%) can be removed *via* the (±)-α-methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* **12** 120 2008.]

**2,6-Dichlorobenzoic acid** [50-30-6] M 191.0, m 141-142°, pK<sup>25</sup> 1.59. Crystallise the acid from EtOH and sublime it *in vacuo*. [*Beilstein* **9** IV 1005.] Aromatic acid impurities (to <0.05%) can be removed *via* the (±)-α-methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* **12** 120 2008.]

**3,4-Dichlorobenzoic acid** [51-44-5] M 191.0, m 206-207°, pK<sup>25</sup> 3.64. Crystallise the acid from aqueous EtOH (charcoal) or acetic acid. [*Beilstein* **9** IV 1006.] Aromatic acid impurities (to <0.05%) can be removed *via* the (±)-α-methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* **12** 120 2008.]

**3,5-Dichlorobenzoic acid** [51-36-5] M 191.0, m 188°, pK<sup>25</sup> 3.54. Crystallise the acid from EtOH and sublime it in a vacuum. [*Beilstein* **9** IV 1008.] Aromatic acid impurities (to <0.05%) can be removed *via* the (±)-α-methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* **12** 120 2008.]

**2,6-Dichlorobenzonitrile** [1194-65-6] M 172.0, m 145°. Crystallise the nitrile from acetone. [*Beilstein* **9** IV 1006.]

**4,4'-Dichlorobenzophenone** [90-98-2] M 251.1, m 145-146°. Recrystallise it from EtOH [Wagner et al. *J Am Chem Soc* **108** 7727 1986]. The *oxime* has m 135.2-136.9° (from MeOH, Sieger & Klein *J Org Chem* **22** 953 1957). The *semicarbazone* has m 192-193° (from H<sub>2</sub>O). [*Beilstein* **7** H 420, **7** I 228, **7** II 359, **7** III 2076, **7** IV 1376.]

**2,5-Dichloro-1,4-benzoquinone** [615-93-0] M 177.0, m 161-162°, 163°. Recrystallise it twice from 95% EtOH to give yellow needles [Beck et al. *J Am Chem Soc* **108** 4018 1986]. The *dioxime* has m 278°(dec). [*Beilstein* **7** H 632, **7** I 346, **7** II 580, **7** III 3376, **7** IV 2081.]

**2,6-Dichloro-1,4-benzoquinone** [697-91-6] M 177.0, m 122-124°. Recrystallise the quinone from petroleum ether (b 60-70°). It sublimes at 41°/17.6μ. [Carlson & Miller *J Am Chem Soc* **107** 479 1985, *Beilstein* **7** II 580, **7** III 3376.]

**2,6-Dichlorobenzoyl chloride** [4659-45-4] M 209.5, m 15-17, b 122-124°/15mm,  $d_4^{20}$  1.464. Reflux the acid chloride for 2 hours with excess of acetyl chloride (3 volumes), distil off AcCl followed by the benzoyl chloride. Store it away from moisture. It is an **IRRITANT**. [Beilstein 9 III 1377.]

**3,4-Dichlorobenzyl alcohol** [1805-32-9] M 177.0, m 38-39°, 148-151°/760mm. Crystallise the alcohol from EtOH (m 32-34°) or water (m 38°, needles). Its solubility at 20° is 1g in 1250ml of H<sub>2</sub>O. [Beilstein 6 H 445, 6 III 1558, 6 IV 2598.]

**2,3-Dichloro-5,6-dicyano-*p*-benzoquinone (DDQ)** [84-58-2] M 227.0, m 203°(dec), 213-216°, (210-215°)(dec). Crystallise DDQ from CHCl<sub>3</sub>, CHCl<sub>3</sub>/\*benzene (4:1), or \*benzene and store it at 0°. [Pataki & Harvey *J Org Chem* 52 2226 1987, Beilstein 10 H 902, 10 II 635, 10 IV 3521.]

**2,4-Dichloro-6-methylphenol (2,4-dichloro-*o*-cresol)** [1570-65-6] M 177.0, m 55°, b 111-113°/16mm, 129-132°/40mm, 231.6-231.9°/760mm, pK<sup>20</sup> 8.14. Crystallise the cresol from Et<sub>2</sub>O (m 55°) or water; it has m 53-54° after sublimation. [Beilstein 6 H 359, 6 II 332, 6 III 1267, 6 IV 2001.]

**2,4-Dichloro-1-naphthol** [2050-76-2] M 213.1, m 106-107°, pK<sub>Est</sub> ~7.7. Crystallise the naphthol from MeOH. The *1-methyl ether* has m 58° (from EtOH). [Beilstein 6 H 612, 6 I 308, 6 II 582, 6 III 2934, 6 IV 4233.]

**2,3-Dichloro-1,4-naphthoquinone** [117-80-6] M 227.1, m 193°. Crystallise the quinone from EtOH. [Beilstein 7 IV 2426.]

**2,5-Dichloro-4-nitroaniline** [6627-34-5] M 207.0, m 153-154°, 157-158°, pK<sup>25</sup> -0.74 (aqueous H<sub>2</sub>SO<sub>4</sub>). Recrystallise it from EtOH, then sublime it *in vacuo*. [Beilstein 12 H 735, 12 II 400.]

**2,6-Dichloro-4-nitroaniline (Dichloran)** [99-30-9] M 207.0, m 193°. Crystallise Dichloran from aqueous EtOH or \*benzene/EtOH. [Beilstein 12 IV 1681.]

**2,5-Dichloro-1-nitrobenzene (1,4-dichloro-4-nitrobenzene)** [89-61-2] M 192.0, m 56°, 266-269°/atm. Crystallise the nitrobenzene from absolute EtOH. [Beilstein 5 IV 726.]

**3,4-Dichloro-1-nitrobenzene (1,2-dichloro-4-nitrobenzene)** [99-54-7] M 192.0, m 43°. Crystallise it from absolute EtOH. [Beilstein 5 IV 726.]

**2,4-Dichloro-6-nitrophenol** [609-89-2] M 208.0, m 122-123°, pK<sub>Est</sub> ~5.0. Crystallise the chloronitrophenol from AcOH. [Beilstein 6 IV 1358.]

**2,6-Dichloro-4-nitrophenol** [618-80-4] M 208.0, m 125°(dec), pK<sup>25</sup> 3.55. Crystallise the chloronitrophenol from EtOH and dry it *in vacuo* over anhydrous MgSO<sub>4</sub>. [Beilstein 6 IV 1361.]

**Dichlorophen [2,2'-methylenebis(4-chlorophenol)]** [97-23-4] M 269.1, b 177-178°, pK<sub>Est</sub> ~9.7. Crystallise dichlorophen from toluene. [Beilstein 6 III 5406.]

**2,3-Dichlorophenol** [576-24-9] M 163.0, m 57°, pK<sup>25</sup> 7.70. Crystallise it from ether. [Beilstein 6 IV 883.]

**2,4-Dichlorophenol** [120-83-2] M 163.0, m 42-43°, pK<sup>25</sup> 7.89. Crystallise it from petroleum ether (b 30-40°). Purify it also by repeated zone melting, using a P<sub>2</sub>O<sub>5</sub> guard tube to exclude moisture. It is very *hygroscopic* when dry. [Beilstein 6 IV 885.]

**2,5-Dichlorophenol** [583-78-8] M 163.0, m 58°, b 211°/744mm, pK<sup>25</sup> 7.51. Crystallise it from ligroin and sublime it in a vacuum or distil it. [Beilstein 6 IV 942.]

**3,4-Dichlorophenol** [95-77-2] M 163.0, m 68°, b 253.5°/767mm, pK<sup>25</sup> 8.58. Crystallise 3,4-dichlorophenol

from petroleum ether/\*benzene mixture and/or distil it. [Beilstein 6 IV 952.]

**3,5-Dichlorophenol** [591-35-5] M 163.0, m 68°, b 122-124°/8mm, 233-234°/760mm, pK<sup>25</sup> 8.81. Crystallise 3,5-dichlorophenol from petroleum ether/\*benzene mixture and/or distil it. [Beilstein 6 IV 957.]

**2,4-Dichlorophenoxyacetic acid (2,4-D)** [94-75-7] M 221.0, m 146°, pK<sup>25</sup> 2.90. Crystallise 2,4-D from MeOH. It is a plant growth substance, a herbicide and is TOXIC. [Beilstein 6 IV 908.]

**(±)-α-(2,4-Dichlorophenoxy)propionic acid (2,4-DP, Dichloroprop)** [120-36-5] M 235.1, m 117°, pK<sup>20</sup> 2.86. Crystallise 2,4-DP from MeOH. It is a plant growth substance, a herbicide and is TOXIC. The *R*(+)- and *S*(-)- enantiomers have m 124° (from \*C<sub>6</sub>H<sub>6</sub>) and [α]<sub>D</sub><sup>25</sup> ±35.2° (c 1, Me<sub>2</sub>CO). [Beilstein 6 H 189, 6 III 708, 6 IV 922-923.]

**2,4-Dichlorophenylacetic acid** [19719-28-9] M 205.0, m 131°, 132-133°, pK<sub>Est</sub> ~4.0. Crystallise the acid from aqueous EtOH. [Beilstein 9 III 2271.]

**2,6-Dichlorophenylacetic acid** [6575-24-2] M 205.0, m 157-158°, pK<sub>Est</sub> ~3.8. Crystallise the acid from aqueous EtOH. [Beilstein 9 III 2272.]

**3-(3,4-Dichlorophenyl)-1,1-dimethyl urea (Diuron)** [330-54-1] M 233.1, m 153-154°, 155°. Recrystallise it from 95% EtOH [Beck et al. *J Am Chem Soc* 108 4018 1986]. [Beilstein 12 IV 1263.]

**4,5-Dichlorophthalic acid** [56962-08-4] M 235.0, m 200° (dec to anhydride), pK<sub>Est(1)</sub> 2.2, pK<sub>Est(2)</sub> ~4.7. Crystallise the acid from water. It has been purified by converting to the anhydride, reacting it with boiling EtOH to form the *monoethyl ester* (m 133-134°), and hydrolysing it back to the diacid (see next entry). [Beilstein 9 III 4205.]

**3,6-Dichlorophthalic anhydride** [4466-59-5] M 189-191°, 191-191.5°, b 339°. Boil the anhydride in xylene (allowing any vapours which would contain H<sub>2</sub>O to be removed, e.g. Dean and Stark trap), which causes any acid present to dehydrate to the anhydride, and cool. Recrystallise it from xylene [Villiger *Chem Ber* 42 3539 1909, Fedorow *Izv Akad Nauk SSSR Otd Khim Nauk* 397 1948, *Chem Abstr* 1585 1948]. [Beilstein 17/11 V 260.]

**2,6-Dichlorostyrene** [28469-92-3] M 173.0, m 8°, b 72-73°/2mm, d<sub>4</sub><sup>20</sup> 1.4045, n<sub>D</sub><sup>20</sup> 1.5798. Purify the styrene by fractional crystallisation from the melt and by distillation *in vacuo*. [Beilstein 5 III 1174.]

**2,4-Dichlorotoluene** [95-73-8] M 161.1, m -13.5°, b 61-62°/3mm, d<sub>4</sub><sup>20</sup> 1.250, n<sub>D</sub><sup>20</sup> 1.5513. Recrystallise 2,4-dichlorotoluene from EtOH at low temperature or fractionally distil it. [Beilstein 5 IV 815.]

**2,6-Dichlorotoluene** [118-69-4] M 161.1, b 199-200°/atm, d<sub>4</sub><sup>20</sup> 1.254, n<sub>D</sub><sup>20</sup> 1.548. Fractionally distil it and collect the middle fraction. [Beilstein 5 IV 815.]

**3,4-Dichlorotoluene** [95-75-0] M 161.1, m -16°, b 205°/atm, d<sub>4</sub><sup>20</sup> 1.2541, n<sub>D</sub><sup>20</sup> 1.549. Recrystallise it from EtOH at very low temperature or fractionally distil it. [Beilstein 5 IV 815.]

**α,α'-Dichloro-*p*-xylene** [623-25-6] M 175.1, m 100°, 254°/atm. Crystallise the xylene from \*benzene and dry it under vacuum. [Beilstein 5 IV 967.]

**Dicinnamalacetone (1,9-diphenyl-1,3,6,8-nonatetraen-5-one)** [622-21-9] M 314.4, m 144°, 146°. Crystallise the ketone from \*benzene/isooctane (1:1). The *1,3,5-trinitrobenzene complex* (1:1) has m 113° (from EtOH), and the *1,3,5-trinitrobenzene complex* (1:2) has m 110° (from EtOH). [Beilstein 7 H 524, 7 I 293, 7 II 484, 7 III 2756, 7 IV 1834.] The *2,4-dinitrophenylhydrazone* has m 199-200°.

**Dicumyl peroxide** [80-43-3] M 270.4, m 39-40°. Crystallise the peroxide from 95% EtOH (charcoal). Store

it at 0°. *Potentially EXPLOSIVE.* [Beilstein 6 IV 3220.]

**9,10-Dicyanoanthracene** [1217-45-4] **M 228.3, m 340°.** Recrystallise the dinitrile twice from pyridine [Mattes & Farid *J Am Chem Soc* 108 7356 1986]. [Beilstein 9 IV 3667.]

**1,4-Dicyanobenzene (terephthalonitrile)** [623-26-7] **M 128.1, m 222°.** Crystallise the dinitrile from EtOH or AcOH, and has **m 221.5-222.5°** after sublimation. [Beilstein 9 H 846, 9 I 376, 9 II 613, 9 III 4255, 9 IV 3328.]

**1,4-Dicyanonaphthalene** [3029-30-9] **M 178.2, m 206°.** Purify it by recrystallisation from EtOH (charcoal), **m 208°**, and sublime it *in vacuo*. [Bradbrook & Linstead *J Chem Soc* 1742 1936, Beilstein 9 H 917, 9 II 651, 9 III 464.]

**trans-4-(Diethylamino)azobenzene** [3588-91-8] **M 320.5, m 171° pK<sup>26</sup> 3.08 (50% aqueous EtOH).** Purify the azobenzene by column chromatography on alumina, elute with, and crystallise it from, toluene, [Albini et al. *J Chem Soc Perkin Trans 2* 1393 1982, Flamigni & Monti *J Phys Chem* 89 3702 1985]. The *deoxycholate (1:4) complex* has **m 194-194°** (from EtOH). [Beilstein 16 H 341, 16 I 311, 16 III 342, 16 IV 455.]

**N,N-Diethylaniline** [91-66-7] **M 149.2, b 216.5°/atm, d<sub>4</sub><sup>20</sup> 0.938, n<sub>D</sub><sup>20</sup> 1.5409 pK<sup>25</sup> 6.57.** Reflux the base for 4 hours with half its weight of acetic anhydride, then fractionally distil it under reduced pressure (**b 92°/10mm**). [Beilstein 12 IV 252.]

**Di-(2-ethylhexyl)phthalate ('di-iso-octyl' phthalate)** [117-81-7] **M 390.6, b 384°, 256-257°/1mm, d<sub>4</sub><sup>20</sup> 0.9803, n<sub>D</sub><sup>20</sup> 1.4863.** Wash the ester with Na<sub>2</sub>CO<sub>3</sub> solution, then shake it with water. After the resulting emulsion has been broken by adding ether, the ethereal solution is washed twice with water, dried (CaCl<sub>2</sub>), and evaporated. The residual liquid is distilled several times under reduced pressure, then stored in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> [French & Singer *J Chem Soc* 1424 1956]. [Beilstein 9 IV 3184.]

**Diethyl phenyl orthoformate (diethoxy phenoxy ethane)** [14444-77-0] **M 196.3, b 111°/11mm, 122°/13mm, d<sub>4</sub><sup>20</sup> 1.0099, n<sub>D</sub><sup>20</sup> 1.4799.** Fractionate the ortho-ester through an efficient column under vacuum [Smith *Acta Chem Scand* 10 1006 1956].

**Diethyl phthalate** [84-66-2] **M 222.2, b 172°/12mm, b 295°/760mm, d<sub>4</sub><sup>25</sup> 1.1160, n 1.5022.** Wash the ester with aqueous Na<sub>2</sub>CO<sub>3</sub>, then distilled water, dry (CaCl<sub>2</sub>), and distil it under reduced pressure. Store it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. [Beilstein 9 IV 3172.]

**Diethyl 2-phthalimidomalonate** [56680-61-5] **M 305.3, m 72-74°, 73-74°, pK<sup>25</sup> 9.17.** Dissolve it in xylene and when the temperature is 30° add petroleum ether (b 40-60°) and cool to 20° whereby the malonate separates as a pale brown powder [Booth et al. *J Chem Soc* 666 1944]. *Alternatively*, dissolve it in \*C<sub>6</sub>H<sub>6</sub>, dry it over CaCl<sub>2</sub>, filter, evaporate and the residual oil solidifies. Grind this with Et<sub>2</sub>O, filter and wash it with Et<sub>2</sub>O until white in colour, and dry it in a vacuum. It forms a yellow *sodium salt* **m 280°(dec)**. The anion has  $\lambda_{\max}$  at 254nm ( $\epsilon$  18.5K) [Clark & Murray *Org Synth Coll Vol I* 271 1941, UV of Na salt: Nnadi & Wang *J Am Chem Soc* 92 4421 1970]. [Beilstein 21 H 487, 21 I 379, 21 III/IV 5264.]

**Diethylstilboesterol [stilbesterol, stilboesterol, (E)-3,4-bis(4-hydroxyphenyl)-3-hexene]** [56-23-1] **M 268.4, m 169-172°.** Crystallise stilbesterol from \*benzene. [Beilstein 6 IV 6856.]

**Diethyl terephthalate** [636-09-0] **M 222.2, m 44°, 142°/2mm, 302°/760mm.** Crystallise the ester from toluene and distil it under reduced pressure. [Beilstein 9 H 8, 9 I 374, 9 III 4250, 9 IV 3304.]

**4,4'-Di-n-heptyloxyazoxybenzene** [2635-26-9] **M 426.6, m 75°, 95° (smectic Ø nematic) and 127° (nematic Ø liquid), pK<sub>Est</sub> ~ -5.** Purify azoxybenzene by chromatography on Al<sub>2</sub>O<sub>3</sub> (\*benzene), recrystallise it from hexane or 95% EtOH and dry it by heating under vacuum. The liquid crystals can be sublimed *in vacuo*. [Mellifiori et al. *Spectrochim Acta Part A* 37(A) 605 1981, Dewar & Schroeder *J Am Chem Soc* 86 5235 1964, Weygand & Glaber *J Prakt Chem* 155 332 1940, Beilstein 16 III 600, 16 IV 5264.]

**9,10-Dihydroanthracene** [613-31-0] M 180.3, m 110-110.5°, b ~312°/atm,  $d_4^{20}$  0.880. Crystallise it from EtOH [Rabideau et al. *J Am Chem Soc* 108 8130 1986]. [Beilstein 5 H 641, 5 IV 2182.]

**Dihydrochloranil (tetrachloro-1,4-hydroquinone)** [87-87-6] M 247.9, m 240.5°. Crystallise the quinone from EtOH or AcOH/EtOH. Sublime it at 77°/0.6x10<sup>-3</sup>mm. The *dibenzoyl* derivative has m 233°. [Conant & Fieser *J Am Chem Soc* 45 2207 1923, Rabideau et al. *J Am Chem Soc* 108 8130 1986, Beilstein 6 H 851, 6 I 417, 6 II 846, 6 III 4436, 6 IV 5775.]

**1,4-Dihydro-1,4-epoxynaphthalene** [573-57-9] M 144.2, m 53-54.5°, 53-56°, 55-56°. Dissolve it in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dry it over K<sub>2</sub>CO<sub>3</sub>, filter, evaporate and dry the residue at 15mm, then recrystallise it from petroleum ether (b 40-60°), dry it at 25°/0.005mm and sublime it (sublimes slowly at room temperature)[Wittig & Pohmer *Chem Ber* 89 1334 1956, Gilman & Gorsich *J Am Chem Soc* 79 2625 1957]. [Beilstein 17 III/IV 548.]

**1,8-Dihydroxyanthraquinone (Danthrone)** [117-10-2] M 240.1, m 193-197°, pK<sub>1</sub><sup>25</sup> 8.30, pK<sub>2</sub><sup>25</sup> 12.46. Crystallise Danthron from EtOH and sublime it in a vacuum. [Beilstein 8 IV 3217.]

**2,4-Dihydroxyazobenzene (Sudan orange G)** [2051-85-6] M 214.2, m 143-146°, 228°, pK<sub>Est(1)</sub> <0, pK<sub>Est(2)</sub> ~7.3, pK<sub>Est(3)</sub> ~9.3. Crystallise the dye from hot EtOH (charcoal). Has UV max at 350nm. [Beilstein 16 IV 264.]

**2,3-Dihydroxybenzaldehyde** [24677-78-9] M 138.1, m 135-136°, pK<sub>1</sub><sup>20</sup> 7.73, pK<sub>2</sub><sup>20</sup> 10.91. Crystallise the aldehyde from water. [Beilstein 8 III 1979.]

**2,4-Dihydroxybenzoic acid (β-resorcylic acid)** [89-86-1] M 154.1, m 226-227°(dec), pK<sub>1</sub><sup>25</sup> 3.30, pK<sub>2</sub><sup>25</sup> 9.12, pK<sub>3</sub><sup>25</sup> 15.6. Crystallise the acid from water. [Beilstein 10 IV 1420.]

**2,5-Dihydroxybenzoic acid (Gentisic acid, 5-hydroxysalicylic acid)** [490-79-9] M 154.1, m 204.5-205°, pK<sup>25</sup> 2.95. Crystallise gentisic acid from hot water or \*benzene/acetone. Dry it in a vacuum desiccator over silica gel. [Beilstein 10 H 384, 10 IV 1441.]

**2,6-Dihydroxybenzoic acid (γ-resorcylic acid)** [303-07-1] M 154.1, m 167°(dec), pK<sup>25</sup> 1.05. Dissolve the acid in aqueous NaHCO<sub>3</sub> and the solution is washed with ether to remove non-acidic material. The acid is precipitated by adding H<sub>2</sub>SO<sub>4</sub>, and recrystallised from water. Dry it under vacuum and store it in the dark [Lowe & Smith *J Chem Soc, Faraday Trans 1* 69 1934 1973]. [Beilstein 10 IV 1456.]

**2,4-Dihydroxybenzophenone** [131-56-6] M 214.2, m 145.5-147° pK<sub>Est(1)</sub> ~7.0, pK<sub>Est(2)</sub> ~12.0. Recrystallise it from MeOH. [Beilstein 8 IV 2442.]

**4,4'-Dihydroxybenzophenone** [611-99-4] M 214.2, m 210°, 213-214°, 213-215°, 216.6-217.1°, pK<sub>Est</sub> ~7.0. The benzophenone was prepared by a Friedel-Crafts reaction between 4-methoxybenzoyl chloride and anisole with anhydrous AlCl<sub>3</sub> in anhydrous CS<sub>2</sub>, followed by demethylation with AlCl<sub>3</sub> in boiling toluene. It could also be made directly from 4-hydroxybenzoic acid and an equivalent of phenol by heating with three parts of freshly fused anhydrous ZnCl<sub>2</sub> at 125-140° for 45 minutes; the cooled mass is treated with dilute HCl, the solid is filtered off, washed thoroughly with aqueous NaHCO<sub>3</sub> solution, H<sub>2</sub>O, dried *in vacuo* and recrystallised from H<sub>2</sub>O, aqueous EtOH, or dilute HCl. [Russell & Butler *J Am Chem Soc* 71 3663 1949.] The *oxime* has m 266-267° (dec) [Zigeuner & Ziegler *Monatsh für Chemie* 89 359 1949], and the *2,4-dinitrophenylhydrazone* has m 190-192° (from EtOH). [Beilstein 8 H 316, 8 I 641, 8 II 355, 8 III 2648, 8 IV 2452.]

**2,5-Dihydroxybenzyl alcohol (Gentisyl alcohol)** [495-08-9] M 140.1, m 47-48°, pK<sub>Est(1)</sub> ~9.3, pK<sub>Est(2)</sub> ~11.3. Crystallise the alcohol from ligroin, CHCl<sub>3</sub>, AcOH or H<sub>2</sub>O. Sublime it at ~70° under high vacuum. [Beilstein 6 II 1084, 6 III 6326.]

**2,2'-Dihydroxybiphenyl** [1806-29-7] M 186.2, m 108.5-109.5°, pK<sub>1</sub><sup>25</sup> 7.56, pK<sub>2</sub><sup>25</sup> 11.80. Crystallise the biphenyl repeatedly from toluene, then sublime it at 60°/10<sup>-4</sup>mm. [Beilstein 6 IV 6645.]

**4,4'-Dihydroxybiphenyl (4,4'-biphenol)** [92-88-6] **M 186.2, m 280.5°, 280-285°, pK<sub>Est(1)</sub> ~3.6, pK<sub>Est(2)</sub> ~11.8.** Recrystallise the biphenol from aqueous EtOH preferably under N<sub>2</sub> to avoid oxidation to the extended quinone. It is characterized as the *dimethyl derivative* (4,4'-dimethoxybiphenyl) from which it is prepared by demethylation. The dimethoxy derivative has **m 176.5-177°** (from AcOH, EtOH, hexane or \*C<sub>6</sub>H<sub>6</sub>) and sublimes *in vacuo*. [Williamson & Rodebush *J Am Chem Soc* **63** 3019 1941, *Beilstein* **6** I 485, **6** II 962, **6** III 6389, **6** IV 6651.]

**1,8-Dihydroxy-3-methylantraquinone (chrysophanic acid)** [481-74-3] **M 245.3, m 196°, pK<sub>Est(1)</sub> ~8.2, pK<sub>Est(2)</sub> ~12.4.** Crystallise chrysophanic acid from EtOH or \*benzene and has **m 195.6-196.2°**, after sublimation it in a vacuum. The yellow *mono-acetate* has **m 188-190°** (from MeOH or Me<sub>2</sub>CO). It forms Ni<sup>2+</sup>, Co<sup>2+</sup> and Cu<sup>2+</sup> complexes. [*Beilstein* **8** H 470, **8** I 725, **8** II 510, **8** III 3808, **8** IV 3277.]

**1,5-Dihydroxynaphthalene (1,5-naphthalenediol)** [83-56-7] **M 160.7, m 258°, 260°, 265°, pK<sub>Est(1)</sub> ~9.0, pK<sub>Est(2)</sub> ~11.0.** The diol (~30g) is purified by making into a thick paste with H<sub>2</sub>O and suspending this in 3L of H<sub>2</sub>O containing 200ml of EtOH, boiling under reflux for 3 hours, cooling to 30°, saturating with SO<sub>2</sub>, digesting below the boiling point for 1 hour and filtering fast through a large hot filter paper. The hot filtrate is poured onto crushed ice whereby the diol (15-20g) separates as colourless needles (**m 258°**) [Wheeler & Ergle *J Am Chem Soc* **52** 4873 1930]. Recrystallise it from H<sub>2</sub>O or nitromethane under N<sub>2</sub> to avoid oxidation. The *dibenzoyl* derivative has **m 245°** (from EtOH). The *5-methoxy-1-naphthol* derivative [prepared from the diol in MeOH/HCl (1:30 weight to volume ratio) and set aside at 25° for 9-10days] crystallised from petroleum ether (**m 135-136°**) or from CH<sub>2</sub>Cl<sub>2</sub>/hexane (needles **m 140°**) [Bell & McCaffrey *Aust J Chem* **46** 731 1993]. [*Beilstein* **6** I 477, **6** II 950, **6** III 5265, **6** IV 6554.]

**1,6-Dihydroxynaphthalene** [575-44-0] **M 160.2, m 138-139° (with previous softening), pK<sub>Est</sub> ~9.4.** Crystallise it from \*benzene or \*benzene/EtOH after treatment with charcoal. [*Beilstein* **6** IV 6557.]

**2,5-Dihydroxyphenylacetic acid (homogentisic acid)** [451-13-8] **M 168.2, m 152°, 154-152°, pK<sup>20</sup> 4.14 (COOH).** Crystallise homogentisic acid from EtOH/CHCl<sub>3</sub> or H<sub>2</sub>O (solubility is 85% at 25°). [*Beilstein* **10** IV 1506.]

**3,4-Dihydroxytoluene (4-methylcatechol)** [452-86-8] **M 124.1, m 65-66°, 68°, b 112°/3mm, 241°/760mm, pK<sub>1</sub><sup>25</sup> 9.44 (9.7), pK<sub>2</sub><sup>25</sup> 10.90 (11.9).** Crystallise the catechol from \*C<sub>6</sub>H<sub>6</sub>. The purity is checked by TLC. Crystallise it from high-boiling petroleum ether and distil it in a vacuum. [*Beilstein* **6** IV 5878.]

**1,4-Diiodobenzene** [624-38-4] **M 329.9, m 132-133°.** Crystallise it from EtOH or boiling MeOH, then dry it in air. [*Beilstein* **5** IV 700.]

**trans-4,4'-Dimethoxyazobenzene** [501-58-6] **M 242.3, m 162.7-164.7°, 165-166°, pK<sub>Est</sub> ~0.** Chromatograph it on basic alumina and elute with \*benzene. Then crystallise the residue from 2:2:1 (v/v) methanol/ethanol/\*benzene or Me<sub>2</sub>CO. [*Beilstein* **16** H 112, **16** I 237, **16** II 43, **16** III 93, **16** IV 172.]

**3,5-Dimethoxybenzaldehyde** [7331-34-4] **M 166.2, m 45.5°, 46-47°, 45-48°, b 151°/16mm.** The aldehyde was prepared by reduction of the corresponding acid chloride with 5%Pd/CaCO<sub>3</sub> in xylene, and was purified by distillation in a vacuum and by recrystallisation from hexane (**m 48°**), ligroin, petroleum ether (**m 45-46°**) or 70% aqueous EtOH. The *oxime* (**m 115°**) crystallises from \*C<sub>6</sub>H<sub>6</sub>, the *thiosemicarbazone* (**m 211-212°**) crystallises from aqueous EtOH, and the *2,4-dinitrophenylhydrazone* (**m 260-261°**) forms orange needles from aqueous EtOH. [Mongolusk et al. *J Chem Soc* 2231 1957, Lambooy *J Am Chem Soc* **76** 133 1954, *Beilstein* **8** II 291, **8** III 2073, **8** IV 1786.]

**1,2-Dimethoxybenzene (veratrole)** [91-16-7] **M 137.2, m 23°, b 208.5-208.7, d<sub>4</sub><sup>20</sup> 1.085, n<sub>D</sub><sup>25</sup> 1.53232.** Steam distil veratrole, then fractionally distil it from BaO, CaH<sub>2</sub> or Na. Crystallise it from \*benzene or low-boiling petroleum ether at 0°. Fractionally crystallise it from its melt. Store it over anhydrous Na<sub>2</sub>SO<sub>4</sub>. [*Beilstein* **6** IV 5564.]

**1,3-Dimethoxybenzene** [151-10-0] **M 137.2, b 212-213°**,  $d_4^{20}$  1.056,  $n_D^{20}$  1.5215. Extract it with aqueous NaOH, and water, then dry it. Fractionally distil it from BaO or Na. [Beilstein 6 IV 5663.]

**1,4-Dimethoxybenzene** [150-78-7] **M 137.2, m 57.2-57.8°**. Steam distil 1,4-dimethoxybenzene, then crystallise it from hexane or \*benzene, and from MeOH or EtOH, but these are wasteful due to high solubilities. Dry it under vacuum. It also sublimes under vacuum. [Beilstein 6 IV 5718.]

**2,4-Dimethoxybenzoic acid** [91-52-1] **M 182.2, m 107.3°**, 109°,  $pK^{25}$  4.36. Crystallise the acid from water and dry it in a vacuum desiccator over  $H_2SO_4$ . The *S*-benzylisothiuronium salt has **m 158-159°** (from  $CHCl_3$ ). [Beilstein 10 H 379, 10 I 177, 10 II 252, 10 III 1371, 10 IV 1422.] Aromatic acid impurities (to <0.05%) can be removed via the ( $\pm$ )- $\alpha$ -methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [Ley & Yates *Organic Process Research & Development* 12 120 2008.]

**2,6-Dimethoxybenzoic acid** [1466-76-8] **M 182.2, m 186-187°**, 188-191°,  $pK^{25}$  3.44. Crystallise the acid from water or 1,1-dichloroethane (**m 187.5-188.5°**). [Beilstein 10 H 388, 10 I 185, 10 II 259, 10 III 1401, 10 IV 1456.]

**3,4-Dimethoxybenzoic acid (veratric acid)** [93-07-2] **M 182.2, m 181-182°**, 185-186°,  $pK^{25}$  4.43. Crystallise the acid from  $Et_2O$ ,  $H_2O$  or aqueous acetic acid. It has **m 180-181°** after sublimation at 80°/1mm. [Beilstein 10 H 393, 10 I 188, 10 II 261, 10 III 1404, 10 IV 1406.]

**3,5-Dimethoxybenzoic acid** [1132-21-4] **M 182.2, m 185-186°**,  $pK^{25}$  3.97. Crystallise the acid from water, EtOH or aqueous acetic acid and dry it in a vacuum. [Beilstein 10 H 405, 10 I 195, 10 III 1446, 10 IV 1501.]

***p,p'*-Dimethoxybenzophenone** [90-96-0] **M 242.3, m 144.5°**. Crystallise the ketone from absolute EtOH, aqueous EtOH or EtOH/AcOH. The 2,4-dinitrophenylhydrazone has **m 199-200°**. [Beilstein 8 H 317, 8 I 641, 8 II 355, 8 III 2649, 8 IV 2453.]

**2,6-Dimethoxy-1,4-benzoquinone** [530-55-2] **M 168.1, m 255-256°**, 256°(dec), 262-263°. Crystallise the quinone from  $H_2O$  or acetic acid. It sublimes at 175-180°/1mm. It has UV with  $\lambda_{max}$  at 287 and 377nm ( $CHCl_3$ ). [Beilstein 8 H 385, 8 I 683, 8 II 433, 8 III 3354, 8 IV 2710.]

**5,6-Dimethoxy-1-indanone** [2107-69-9] **M 192.2, m 118-120°**. Crystallise the indanone from MeOH, then sublime it in a vacuum. [Beilstein 8 IV 1985.]

**1,4-Dimethoxynaphthalene** [10075-62-4] **M 188.2, m 87-88°**. Crystallise the naphthalene from EtOH, MeOH (**m 85-86°**) or petroleum ether. [Beilstein 6 H 984, 6 III 5261, 6 IV 6546.]

**1,5-Dimethoxynaphthalene** [10075-63-5] **M 188.2, m 183-184°**, **b 179°/13mm**. Crystallise it from EtOH, AcOH (**m 178-180°**) or  $*C_6H_6$ . Also distil it in a vacuum. [Beilstein 6 III 5266, 6 IV 6554.]

**3,4-Dimethoxy-6-nitrobenzaldehyde (6-nitroverataldehyde)** [20357-25-9] **M 211.2, m 132°**, 133.5-134.5°. The aldehyde is purified by dissolving 9g in 200ml of boiling 95% EtOH, and set aside overnight to crystallise. It is then dried *in vacuo* at 50° and recrystallised from 110ml of 95% EtOH to give 6-7g of aldehyde with **m 132-133°**. Crystallisation from aqueous EtOH provides light yellow needles. It is light sensitive and should be stored in the dark [Fetscher *Org Synth Coll Vol* 4 735 1963]. [Beilstein 8 H 262, 8 I 610, 8 II 290, 6 III 2065, 6 IV 1785.]

**2,6-Dimethoxyphenol (pyrogallol-1,3-diethylether)** [91-10-1] **M 154.2, m 54-56°**,  $pK_{Est}$  ~9.6. Purify the phenol by zone melting or sublimation in a vacuum. [Beilstein 6 IV 7329.]

**3,5-dimethoxyphenol (phloroglucinol dimethylether)** [500-99-2] **M 154.2, m 42-43°**, **b 115°/0.04mm**,  $pK^{25}$  9.35. Purify the phenol by distillation followed by sublimation in a vacuum. [Beilstein 6 IV 7362.]

**3,4-Dimethoxyphenyl acetic acid (homoveratric acid)** [93-40-3] **M 196.2, m 97-99°, pK<sup>25</sup> 4.33.** Crystallise homoveratric acid from H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>/ligroin. The *amide* has **m 142°** (from H<sub>2</sub>O). [*Beilstein* **10** H 409, **10** I 197, **10** II 268, **10** III 1459, **10** IV 1509.]

**3,5-Dimethoxyphenylacetonitrile** [13388-75-5] **M 177.1, m 53°.** Crystallise the nitrile from MeOH or petroleum ether (b 90-110°). [Adams et al. *J Am Chem Soc* **70** 664 1948, Sankaraman et al. *J Am Chem Soc* **109** 5235 1987, *Beilstein* **10** I 198, **10** II 269, **10** III 1470.]

**4,4'-Dimethoxythiobenzophenone** [958-80-5] **M 258.3, m 120°.** Recrystallise the thioketone from a mixture of cyclohexane/dichloromethane (4:1) or EtOH (**m 119°**). [Bergmann & Wagenberg *Chem Ber* **63** 2590 1930, *Beilstein* **8** H 319, **8** II 355, **8** III 2658, **8** IV 2457.]

**2,6-Dimethoxytoluene** [5673-07-4] **M 152.2, m 39-41°, b 97-99°/15mm, 219-220°/731mm.** Sublime 2,6-dimethoxytoluene *in vacuo*. Distil it (preferably under reduced pressure) and/or recrystallise it from pentane, EtOH or aqueous MeOH. [Sankaraman et al. *J Am Chem Soc* **109** 5235 1987]. [*Beilstein* **6** H 872, **6** III 4513, **6** IV 5877.]

**4,4'-Dimethoxytrityl chloride (DMT)** [40615-36-9] **M 338.8, m 114°.** DMT crystallises from cyclohexane/acetyl chloride as the hydrochloride. Dry it over KOH pellets in a desiccator. When dissolved in \*C<sub>6</sub>H<sub>6</sub> and air is blown through, HCl is removed. It crystallises from Et<sub>2</sub>O. [Baeyer & Villiger *Chem Ber* **36** 2788 1903, Smith et al. *J Am Chem Soc* **84** 430 1962, Smith et al. *J Am Chem Soc* **85** 3821 1963.] If it has hydrolysed considerably (see OH in IR), then repeat the crystallisation from cyclohexane/acetyl chloride — excess of AcCl is removed in a vacuum over KOH, then recrystallise it from Et<sub>2</sub>O. [*Beilstein* **6** IV 1042.]

***p*-Dimethylaminoazobenzene (Dimethyl Yellow)** [60-11-7] **M 225.3, m 118-119°(dec), pK<sub>1</sub><sup>25</sup> -5.34 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>25</sup> 2.96.** Crystallise the dye from acetic acid or isooctane, or from 95% EtOH by adding hot water and cooling. Dry it over KOH under vacuum at 50°. [*Beilstein* **6** IV 448.] **CARCINOGEN.**

**4-*N,N'*-Dimethylaminoazobenzene-4'-isothiocyanate {DABITC, 4-[(4-isocyanatophenyl)-azo]-*N,N'*-dimethylaniline}** [7612-98-8] **M 282.4, m 170-171°, pK<sub>Est</sub> ~2.5.** Crystallise DABITC by dissolving 1g in 150ml of boiling Me<sub>2</sub>CO, filtering hot and allowing to cool at -20° overnight, collecting the solid and drying it in a vacuum. Solutions in pyridine should be used immediately, otherwise it decomposes. It is moisture sensitive. [Chang *Methods Enzymol* **91** 79, 455 1983.]

***p*-Dimethylaminobenzaldehyde (Ehrlich's Reagent, DMAB)** [100-10-7] **M 149.2, m 74-75°, pK<sub>Est</sub> ~2.6.** Crystallise DMAB from water, hexane, or from EtOH (2ml/g), after charcoal treatment, by adding excess of water. *Alternatively*, dissolve it in aqueous acetic acid, filter, and precipitate it with ammonia. Finally recrystallise it from EtOH. It is used for the detection of pyrroles [Iyer et al. *J Org Chem* **59** 6038 1994]. [*Beilstein* **14** IV 51.]

***p*-Dimethylaminobenzoic acid** [619-84-1] **M 165.2, m 242.5-243.5°(dec), pK<sub>1</sub> 2.51, pK<sub>2</sub> 6.03.** Crystallise the acid from EtOH/water. [*Beilstein* **14** IV 1164.]

***p*-Dimethylaminobenzophenone** [530-44-9] **M 225.3, m 92-93°, pK<sub>Est</sub> ~2.7.** Crystallise the pale green *p*-dimethylaminobenzophenone from EtOH. Dissolve 100g in 600ml of boiling EtOH, add 5g of charcoal, cool, isolate the solid by centrifugation and similarly wash the pale crystals with ice-cold EtOH. When filtered by suction, EtOH solution remains on the crystals and turns deep green in air. Dry it in a vacuum and store it in the dark. The *hydrazone* has **m 128-130°** and forms a ketyl with potassium. [Hurd & Webb *Org Synth Coll Vol I* 217 1941, *Beilstein* **14** H 82, **14** I 288, **14** III 218, **14** IV 248.]

***N,N*-Dimethylamino-*p*-chlorobenzene (*p*-chloro-*N,N*-dimethylaniline)** [698-69-1] **M 155.6, m 32-33.5°, 35.5°, b 231°/atm.** Purify it by vacuum sublimation [Guarr et al. *J Am Chem Soc* **107** 5104 1985]. The *picrate* has **m 126-128°** (from methanol).

**4-Dimethylamino cinnamaldehyde** [6203-18-5] **M 175.2, m 141°**. The aldehyde crystallises from EtOH or ligroin and is dried *in vacuo*. The *oxime* has **m 157°** (from ligroin). The *phenylhydrazone* has **m 169°** (from MeOH). It is used as a reagent for amines (with NH<sub>3</sub>, UV has  $\lambda_{\max}$  at 630nm). [König et al. *Chem Ber* **61** 2075 2075, Quareshi & Kahn *Anal Chim Acta* **86** 309 1976, *Beilstein* **14** III 184, **14** IV 197.]

**4-Dimethylamino cinnamic acid** [1552-96-1] **M 191.2, m 225° (effervesces), pK<sub>Est(1)</sub> ~2.2, pK<sub>Est(2)</sub> ~4.6**. Recrystallise the acid from EtOH. The *methyl ester* has **m 134-135°** (from Me<sub>2</sub>CO), and the *ethyl ester* has **m 77-78°** (from aqueous EtOH). [Shopper *J Chem Soc* 982 1938, Galat *J Am Chem Soc* **68** 376 1946, *Beilstein* **14** H 522, **14** II 318, **14** III 1306, **14** IV 1716.]

**2S,3R-(+)-4-Dimethylamino-1,2-diphenyl-3-methyl-2-butanol** [38345-66-3] **M 283.4, m 55-57°, [α]<sub>D</sub><sup>20</sup> +9.3° (c 9.6, EtOH), [α]<sub>D</sub><sup>20</sup> +7.7° (c 9.6, EtOH), pK<sub>Est</sub> ~10.0**. Purify the *hydrochloride* by dissolving 1.5g in 13.5 ml of 5N HCl, heat to boiling and evaporate in a vacuum. Recrystallise the *hydrochloride* three times from MeOH/EtOAc giving **m 189-190°, [α]<sub>D</sub> -33.7° (c 1, H<sub>2</sub>O)** {enantiomer has +34.2°}. The *hydrochloride* in the minimum volume of water is basified with aqueous 5N NaOH and extracted with Et<sub>2</sub>O. The extract is dried (K<sub>2</sub>CO<sub>3</sub>) and evaporated, leaving a residue which is stored in a desiccator over solid KOH as a low melting solid. It can be recovered using these procedures from asymmetric reductions with LAH, and re-used. [Pohland & Sullivan *J Am Chem Soc* **77** 3400 1955, Sullivan et al. *J Org Chem* **28** 2483 1963, *Beilstein* **13** IV 2221.]

**N,N-Dimethylaniline** [121-69-7] **M 121.2, f 2°, b 84°/15mm, 193°/760mm, d<sub>4</sub><sup>20</sup> 0.956, n<sub>D</sub><sup>25</sup> 1.5556, pK<sup>25</sup> 5.07**. Primary and secondary amines (including aniline and monomethylaniline) can be removed by refluxing for 4-5 hours with excess acetic anhydride, and then fractionally distilling. Crocker and Jones (*J Chem Soc* 1808 1959) used four volumes of acetic anhydride, then distilled off the greater part of it, and dissolved the residue in ice-cold dilute HCl. Non-basic materials were removed by ether extraction, then the dimethylaniline was liberated with ammonia, extracted with ether, dried, and distilled under reduced pressure. Metzler and Tobolsky (*J Am Chem Soc* **76** 5178 1954) refluxed with only 10% (w/w) of acetic anhydride, then cooled and poured it into excess 20% HCl, which, after cooling, was extracted with diethyl ether. (The amine hydrochloride remains in the aqueous phase.) The HCl solution was cautiously made alkaline to phenolphthalein, and the amine layer was drawn off, dried over KOH and fractionally distilled under reduced pressure, under nitrogen. Suitable drying agents for dimethylaniline include NaOH, BaO, CaSO<sub>4</sub>, and CaH<sub>2</sub>. Other purification procedures include the formation of the *picrate* (**m 163°** from Me<sub>2</sub>CO or EtOH/H<sub>2</sub>O), prepared in \*benzene solution and crystallised to constant melting point, then decomposed with warm 10% NaOH and extracted into ether: the extract was washed with water and distilled under reduced pressure. The *oxalate* salt has also been used for purification. The base has been fractionally crystallised by partial freezing and also from aqueous 80% EtOH then from absolute EtOH. It has been distilled from zinc dust, under nitrogen. [*Beilstein* **12** H 141, **12** I 151, **12** II 2, **12** III 245, **12** IV 243.]

**2,3-Dimethylaniline** [2,3-xylidine (*vic, o*)] [87-59-2] **M 121.2, m 2°, b 106°/15mm, 223°/760mm, d<sub>4</sub><sup>20</sup> 1.570, n<sub>D</sub><sup>20</sup> 0.991, pK<sup>25</sup> 4.70**. Purify *vic*-xylidine by conversion into a derivative (see below), recrystallise the derivative, decompose the derivative with aqueous NaOH and fractionally distil the liquid base. The *acetyl* derivative has **m 135°** (from EtOH), and the *formyl* derivative has **m 102°** (from EtOH). [*Beilstein* **12** H 1101, **12** III 2438, **12** IV 2497.]

**2,4-Dimethylaniline** [2,4-xylidine (*uns, m*)] [95-68-1] **M 121.2, m 11°, b 212°/736mm, 213-214°/760mm, d<sub>4</sub><sup>20</sup> 0.974, n<sub>D</sub><sup>20</sup> 1.5604, pK<sup>25</sup> 4.89**. Convert *uns*-xylidine to a derivative (see below) which, after recrystallisation, is decomposed with alkali to give the free base. Dry it over KOH and fractionally distil. The *acetyl* derivative has **m 130°**, the *benzoyl* derivative has **m 192°**, and the *picrate* has **m 209°**. [*Beilstein* **12** H 1111, **12** IV 2545.]

**2,5-Dimethylaniline** [2,5-xylidine (*p*)] [95-78-3] **M 121.2, m 11°, 15.5°, b 104°/15mm, 215°/739mm, 218°/760mm, d<sub>4</sub><sup>20</sup> 0.974, n<sub>D</sub><sup>20</sup> 1.5604, pK<sup>25</sup> 4.53**. Convert *p*-xylidine to a derivative (see below) which, after recrystallisation, is decomposed with alkali to give the free base. Dry over KOH and fractionally distil. The *acetyl* derivative has **m 142°** (from H<sub>2</sub>O or toluene), and the *benzoyl* derivative has **m 140°** (from EtOH). [*Beilstein* **12** H 1135, **12** IV 2567.]

**2,6-Dimethylaniline** [2,6-xylylidine (*vic, m*)] [87-62-7] **M 121.2, m 11°, b 98°/14mm, 210-211°/736mm, 215°/760mm,  $d_4^{20}$  0.974,  $n_D^{20}$  1.5604, pK<sup>25</sup> 3.95.** Convert *vic*-xylylidine to a derivative (see below) which, after recrystallisation, is decomposed with alkali to give the free base. Dry it over KOH and fractionally distil. The *formyl derivative* [*N*-(2,6-dimethylphenyl)formamide] **M 149.2**, [607-92-1] has **m 168-170°**, the *acetyl derivative* has **m 177°**, the *benzoyl derivative* has **m 168°**, and the *picrate* has **m 180°**. [Beilstein 12 H 1107, 12 I 482, 12 II 604, 12 III 2462, 12 IV 2521.]

**3,4-Dimethylaniline** [3,4-xylylidine] [95-64-7] **M 121.2, m 51°, b 116-118°/25mm, b 226°/760mm, pK<sup>25</sup> 5.17.** Crystallise it from ligroin and distil it under vacuum. [Beilstein 12 H 1103, 12 IV 2502.]

**3,5-Dimethylaniline** [3,5-xylylidine (*sym, m*)] [108-69-0] **M 121.2, b 105-106°/15mm, 221-222°/atm,  $d_4^{20}$  0.974,  $n_D^{20}$  1.557, pK<sup>25</sup> 4.91.** Convert *sym*-xylylidine to a derivative (see below) which, after recrystallisation, is decomposed with alkali to give the free base. Dry it over KOH and fractionally distil it. The *acetyl derivative* has **m 144°**, the *benzoyl derivative* has **m 136°** and the *picrate* has **m 209°** (from H<sub>2</sub>O, EtOH or 10% AcOH). [Beilstein 12 H 1131, 12 IV 2561.]

**9,10-Dimethylantracene** [781-43-1] **M 206.3, m 180-181°.** Purify 9,10-dimethylantracene by crystallising from EtOH, and by recrystallising from the melt. [Beilstein 5 IV 2329.]

**7,12-Dimethylbenz[*a*]anthracene (DMBA)** [57-97-6] **M 256.4, m 122-123°.** Purify DMBA by chromatography on alumina/toluene or \*benzene. Crystallise it from acetone/EtOH. [Beilstein 5 IV 2587.]

**2,3-Dimethylbenzoic acid** [603-79-2] **M 150.2, m 146°, pK<sup>25</sup> 3.72.** Crystallise the acid from EtOH. It is volatile in steam. The *amide* has **m 156°** (from H<sub>2</sub>O). [Beilstein 9 H 531, 9 III 2434, 9 IV 1797.]

**2,4-Dimethylbenzoic acid** [611-01-8] **M 150.2, m 125-126°, 126-127°, b 267°/727mm, pK<sup>25</sup> 4.22.** Crystallise the acid from EtOH or H<sub>2</sub>O, and sublime it in a vacuum. [Beilstein 9 H 531, 9 III 2436, 9 IV 1801.]

**2,5-Dimethylbenzoic acid** [610-72-0] **M 150.2, m 134°, b 268°/760mm, pK<sup>25</sup> 4.00.** Steam distil the acid, then crystallise it from EtOH or H<sub>2</sub>O (**m 134-134.5°**). [Beilstein 9 H 534, 9 IV 1802.]

**2,6-Dimethylbenzoic acid** [632-46-2] **M 150.2, m 117°, b 118-119°/2.5mm, pK<sup>25</sup> 3.35.** Steam distil the acid, and crystallise it from EtOH or H<sub>2</sub>O (**m 116.3-116.7°**). The *N*-dimethylamide has **m 62-63°** (from Et<sub>2</sub>O). [Beilstein 9 H 531, 9 IV 1798.]

**3,4-Dimethylbenzoic acid** [619-04-5] **M 150.2, m 166°, pK<sup>25</sup> 4.50.** Crystallise it from EtOH or H<sub>2</sub>O (**m 168-168.5°**), and sublime it *in vacuo*. The *phenyl ester* has **m 68°** (from EtOH or petroleum ether) and **b 155-157°/2mm**. [Beilstein 9 II 353, 9 III 2441, 9 IV 1803.]

**3,5-Dimethylbenzoic acid** [499-06-9] **M 150.2, m 170°, pK<sup>25</sup> 4.30.** Distil the acid in steam, crystallise it from H<sub>2</sub>O (**m 171.2-171.7°**) or EtOH, and sublime it *in vacuo*. [Beilstein 9 H 536, 9 III 244, 9 IV 1806.]

**4,4'-Dimethylbenzophenone** [611-97-2] **M 210.3, m 95°, b 150-152°/2mm, 333-334°/725mm.** Purify the benzophenone by zone refining or distillation, preferably in a vacuum. [Beilstein 7 III 2181, 7 IV 1434.]

**2,5-Dimethyl-1,4-benzoquinone** [137-18-8] **M 136.1, m 124-125°, 126-127°.** Crystallise the quinone from \*EtOH. [Beilstein 7 IV 2090.]

**2,6-Dimethyl-1,4-benzoquinone (Phloron)** [527-61-7] **M 136.1, m 72° (sealed tube).** Crystallise the quinone from water/EtOH (8:1). [Beilstein 7 IV 657, 7 III 3402, 7 IV 2096.]

**3,3'-Dimethylcarbanilide (*N,N'*-bis[*m*-tolyl]urea)** [620-50-8] **M 240.3, m 220°, 225°.** Crystallise urea from ethyl acetate. The UV has  $\lambda_{\max}$  at 258nm (EtOH). The *hydrochloride* has **m 162°**. [Beilstein 12 III 1970, 12 IV 1829.]

**1,1-Dimethyl-1H-indene** [18636-55-0] **M 144.2, b 57°/4.8mm, 115°/20mm.** Purify the oily indene by gas chromatography or by fractional distillation. [Bosch & Brown *Can J Chem* **42** 1718 1964, cf. *Beilstein* **5** I 251, **5** III 1377.]

**1,5-Dimethylnaphthalene** [571-61-9] **M 156.2, m 81-82°, b 265-266°.** Crystallise it from 85% aqueous EtOH. [*Beilstein* **5** IV 1709.]

**2,3-Dimethylnaphthalene** [581-40-8] **M 156.2, m 104-104.5°.** Steam distil the naphthalene and crystallise it from EtOH. [*Beilstein* **5** IV 1713.]

**2,6-Dimethylnaphthalene** [581-42-0] **M 156.2, m 110-111°, b 122.5-123.5°/10mm, 261-262°/760mm.** Distil it in steam and crystallise it from EtOH. [*Beilstein* **5** IV 1714.]

**3,3'-Dimethylnaphthidine (4,4'-diamino-3,3'-dimethyl-1,1'-binaphthyl)** [13138-48-2] **M 312.4, m 213°.** Recrystallise the naphthidine from EtOH or petroleum ether (b 60-80°). [*Beilstein* **13** IV 493.]

**N,N-Dimethyl-m-nitroaniline** [619-31-8] **M 166.1, m 60°, 61°, pK<sup>25</sup> 2.63.** Crystallise the aniline from EtOH. The *picrate* has **m 119°** (from EtOH). [*Beilstein* **12** H 701, **12** III 1544, **12** IV 1591.]

**N,N-Dimethyl-p-nitroaniline** [100-23-2] **M 166.1, m 164.5-165.2°, pK<sup>25</sup> 0.61 (0.92).** Crystallise the nitroaniline from aqueous EtOH, EtOH or MeOH (**m 163.5-164°**). Dry it *in vacuo*. The *N-methiodide* has **m 161°(dec)** (from H<sub>2</sub>O). [*Beilstein* **12** H 714, **12** III 1584, **12** IV 1616.]

**N,N-Dimethyl-p-nitrosoaniline (4-nitroso-N,N-dimethylaniline)** [138-89-6] **M 150.2, m 86-87°, 92.5-93.5°, b 191-192°/100mm, pK<sup>25</sup> 4.54.** Recrystallise the nitroso-aniline from petroleum ether or CHCl<sub>3</sub>/CCl<sub>4</sub> and dry it in air. *Alternatively*, suspend it in H<sub>2</sub>O, heat to boiling and add HCl until it dissolves. Filter, cool and collect the *hydrochloride* [42344-05-8] with **m 177°** after recrystallisation from H<sub>2</sub>O containing a small amount of HCl. The *hydrochloride* (e.g. 30g) is made into a paste with H<sub>2</sub>O (100ml) in a separating funnel. Add cold aqueous 2.5 NaOH or Na<sub>2</sub>CO<sub>3</sub> to a pH of ~8.0 (green color due to the free base) and extract with toluene, CHCl<sub>3</sub> or Et<sub>2</sub>O. Dry the extract (K<sub>2</sub>CO<sub>3</sub>), filter, distil off the solvent, cool the residue and collect the crystalline free base. Recrystallise it as above and dry it in air. [*Beilstein* **12** IV 1558.]

**R-(+)-N,N'-Dimethyl-1-phenethylamine** [19342-01-9] **and S-(-)-N,N'-Dimethyl-1-phenethyl-amine** [17279-31-1] **M 149.2, b 81°/16mm, [α]<sub>D</sub><sup>20</sup> (+) and (-) 50.2° (c 1, MeOH), d<sub>4</sub><sup>20</sup> 0.908, pK<sub>Est</sub> ~9.0 (for RS).** The amine is mixed with aqueous 10N NaOH and extracted with toluene. The extract is washed with saturated aqueous NaCl, dried over K<sub>2</sub>CO<sub>3</sub>, and transferred to fresh K<sub>2</sub>CO<sub>3</sub> until the solution is clear, and is filtered. The filtrate is distilled. If a short column packed with glass helices is used, the yield is reduced but a purer product is obtained. [Ingersoll *Org Synth* **25** 89 1945, Snyder & Brewster *J Am Chem Soc* **71** 291 4165 1949, Cope et al. *J Am Chem Soc* **71** 3931 1949.] The (-)-*picrate* has **m 140-141°** (from EtOH). The *racemate* [1126-71-2] has **b 88-89°/16mm, 92-94°/30mm, 194-195°/760mm, d<sub>4</sub><sup>20</sup> 0.908.** [*Beilstein* **13** III 2392.]

**2,3-Dimethylphenol** [526-75-0] **M 122.2, m 75°, b 120°/20mm, 218°/760mm, pK<sup>25</sup> 10.54.** Crystallise 2,3-xyleneol from aqueous EtOH. [*Beilstein* **6** IV 3096.]

**2,5-Dimethylphenol** [95-87-4] **M 122.2, m 73°, b 211.5°/762mm, pK<sup>25</sup> 10.41.** Crystallise 2,5-xyleneol from EtOH/ether. [*Beilstein* **6** IV 3164.]

**2,6-Dimethylphenol** [576-26-1] **M 122.2, m 49°, b 203°/760mm, pK<sup>25</sup> 10.61.** Fractionally distil 2,6-xyleneol under nitrogen, crystallise it from \*benzene or hexane, and sublime it at 38°/10mm. [*Beilstein* **6** IV 3122.]

**3,4-Dimethylphenol** [95-65-8] **M 122.2, m 62.5°, 65°, 65-68°, 67°, b 142°/50mm, 225°/757mm, pK<sup>25</sup> 10.36.** Heat 3,4-xyleneol with an equal weight of conc H<sub>2</sub>SO<sub>4</sub> at 103-105° for 2-3 hours, then dilute it with four volumes of water, reflux it for 1 hour, and either steam distil or extract it repeatedly with diethyl ether after cooling to room temperature. The steam distillate is also extracted and evaporated to dryness. (The purification

process depends on the much slower sulfonation of 3,4-dimethylphenol than most of its likely contaminants.) It can also be crystallised from water, hexane or petroleum ether, and sublimed in a vacuum. [Kester *Ind Eng Chem (Anal Ed)* **24** 770 1932, Bernasconi & Paschalis *J Am Chem Soc* **108** 2969 1986, *Beilstein* **6** IV 3099.]

**3,5-Dimethylphenol** [108-68-9] **M 122.2, m 68°, b 219°, pK<sup>25</sup> 10.19.** Purify it as for 3,4-dimethylphenol. [*Beilstein* **6** IV 3141.]

**2,6-Dimethylphenylisocyanide** [2769-71-3] **M 131.2, m 72-76°, 72-75°.** Like most isocyanides (isonitriles, carbylamines), it reacts with transition metals to form aldimines [Malatesta & Bonati *Isonitrile Complexes of Metals* Wiley Interscience, NY 1969, SBN 470 565551, Treichel in *Advances in Organometallic Chemistry* **Vol 11**, (Stone & West Eds) Academic Press NY 1973], and because of the crowded area around the nitrogen atom it coordinates with the metal in metal catalysts to enhance their stereoselectivity. The isocyanide can function as a nucleophile, electrophile, carbene, radical acceptor or pseudohalogen [Ugi (ed.) *Isonitrile Chemistry Academic Press NY 1971*, Library of Congress No 73-84156]. Among the general methods for preparing isocyanides the one that uses the *Vilsmeier reagent* {(chloromethylene)dimethyliminium chloride, [ClCH=NMe<sub>2</sub>]<sup>+</sup> Cl<sup>-</sup>, **M 128.0, m 132°(dec)**, [3724-43-4]} has been commonly used, and in this case it is prepared *in situ*. Thus a solution of redistilled SOCl<sub>2</sub> (0.5mol) in DMF (150ml) is added slowly to a stirred solution of \*N-(2,6-dimethylphenyl) formamide (0.5mol, see below) {**M 149.2, m 165°**, see [607-92-1]} under N<sub>2</sub> and the temperature is kept below -50°. On completion, the cooling bath is momentarily removed to allow the temperature to rise to -45°, then the cooling bath is replaced while anhydrous Na<sub>2</sub>CO<sub>3</sub> (1mol) is added very carefully with rapid stirring till evolution of CO<sub>2</sub> ceases. Stirring is continued while the mixture is heated to 35°, then kept at ~25° for 1 hour, diluted with ice-water and extracted into pentane. The extract is separated, dried (Na<sub>2</sub>SO<sub>4</sub>), and evaporated to give 2,6-dimethylphenylisocyanide as a crystalline colourless solid in 74% yield. It can be recrystallised from pentane, and should be stored under N<sub>2</sub> between 2° to 8°. It is **toxic**. [Periasamy and Walborsky *Org Prep Proc Int* **11** 295 1979.]

\*N-(2,6-Dimethylphenyl)formamide, [607-92-1] **M 149.2, m 164° (164-5°, 168-170°** also reported, and **176-177°** on rapid heating) is prepared by heating 2,6-dimethylaniline (see [87-62-7]) with formic acid, pouring into H<sub>2</sub>O, filtering, drying and recrystallising from EtOH. [Hodgkinson & Limpach *J Chem Soc* **77** **65** (67) 1900, *Beilstein* **12** H 1109, **12** I 604.]

**Dimethyl phthalate** [131-11-3] **M 194.2, b 282°/760mm, d<sub>4</sub><sup>20</sup> 1.190, d<sub>4</sub><sup>25</sup> 1.1865, n<sub>D</sub><sup>20</sup> 1.5149.** Wash the ester with aqueous Na<sub>2</sub>CO<sub>3</sub>, then distilled water, dry (CaCl<sub>2</sub>) and distil it under reduced pressure (**b** 151-152°/0.1mm). [*Beilstein* **7** IV 3170.]

**2,4-Dimethylresorcinol** [634-65-1] **M 138.1, m 111°, 112°, pK<sub>Est(1)</sub> ~9.8, pK<sub>Est(2)</sub> ~11.7.** Crystallise the resorcinol from petroleum ether (**b** 60-80°), \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or toluene/BuOAc. It sublimes at 100-110°/1mm. [*Beilstein* **6** H 918, **6** II 891, **6** III 4588, **6** IV 5955.]

**Dimethyl terephthalate** [120-61-6] **M 194.2, m 141.5-141.8°, 142°.** Purify it by recrystallisation from aqueous EtOH, MeOH or CCl<sub>4</sub>; or by zone melting. [*Beilstein* **6** H 843, **6** III 4250, **6** IV 3303.]

**N,N-Dimethyl-o-toluidine** [609-72-3] **M 135.2, b 68°/10mm, 211-211.5°/760mm, d<sub>4</sub><sup>20</sup> 0.937, n<sub>D</sub><sup>20</sup> 1.53664, pK<sup>25</sup> 6.11.** Isomers and other bases are removed by heating in a water bath for 100 hours with two equivalents of 20% HCl and two and a half volumes of 40% aqueous formaldehyde, then making the solution alkaline and separating the free base. After washing well with water, it is distilled at 10mm pressure and redistilled at ambient pressure [von Braun & Aust *Chem Ber* **47** 260 1914]. Other procedures include drying with NaOH, distilling from zinc in an atmosphere of nitrogen under reduced pressure, and refluxing with excess of acetic anhydride in the presence of conc H<sub>2</sub>SO<sub>4</sub> as catalyst, followed by fractional distillation in a vacuum. The *picrate* has **m** 124-125°. [*Beilstein* **12** H 857, **12** III 1843, **12** IV 1747.]

**N,N-Dimethyl-m-toluidine (m-methyl-N,N-dimethylaniline)** [121-72-2] **M 135.2, b 72-74°/5mm, 128°/57mm, 211.5-212.5°/760mm, d<sub>4</sub><sup>20</sup> 0.93, n<sub>D</sub><sup>20</sup> 1.5500, pK<sup>25</sup> 5.34.** Reflux it for 3 hours with 2 molar equivalents of Ac<sub>2</sub>O, then fractionally distil it under reduced pressure. *Alternatively*, dry over BaO, distil and store it over KOH. The *hydrochloride* has **m** 176° (from EtOH) and the *picrate* has **m** 131°. Methods described

for *N,N*-dimethylaniline are applicable. [Beilstein 12 H 857, 12 III 1953, 12 IV 1815.]

*N,N*-Dimethyl-*p*-toluidine (*p*-methyl-*N,N*-dimethylaniline) [99-97-3] M 135.2, b 76.5-77.5°/4mm, 93-94°/11mm, b 211°/760mm,  $d_4^{20}$  0.937,  $n_D^{20}$  1.5469,  $pK^{21.5}$  5.56,  $pK^{25}$  5.63. Reflux for 3 hours with 2 molar equivalents of  $Ac_2O$ , then fractionally distill it under reduced pressure. Alternatively, dry it over  $BaO$ , distill and store it over  $KOH$ . The picrate has m 128° (from EtOH). Methods described for *N,N*-dimethylaniline are applicable here. [Beilstein 12 H 902, 12 III 2026, 12 IV 1874.]

2,4-Dinitroaniline [97-02-9] M 183.1, m 180°,  $\epsilon_{348}$  12,300 in dilute aqueous  $HClO_4$ ,  $pK^{25}$  -4.27 (aqueous  $H_2SO_4$ ). Crystallise the nitroaniline from boiling EtOH by adding one-third volume of  $H_2O$  and cooling slowly. Dry it on a steam oven. The *N*-acetyl derivative has m 180°. [Beilstein 12 IV 1689.]

2,6-Dinitroaniline [606-23-4] M 183.1, m 139-140°,  $pK^{25}$  -5.37 (aqueous  $H_2SO_4$ ). Purify the nitroaniline by chromatography on alumina, then crystallise it from \*benzene or EtOH. The *N*-acetyl derivative has m 197° (from EtOH). [Beilstein 12 IV 1729.]

2,4-Dinitroanisole [5327-44-6] M 198.1, m 87-88°(metastable), 94-95°, m 206-207°/12mm. Crystallise the anisole from aqueous EtOH. The naphthalene complex has m 50° (from EtOH). [Beilstein 6 H 254, 6 III 858, 6 IV 1372.]

3,5-Dinitroanisole [119-27-7] M 198.1, m 105-106°. Purify the anisole by repeated crystallisation from EtOH, MeOH or  $H_2O$  and dry it in a vacuum desiccator over  $P_2O_5$ . The naphthalene complex has m 69° (from EtOH). [Beilstein 6 III 869, 6 IV 1385.]

1,2-Dinitrobenzene [528-29-0] M 168.1, m 116.5°. Crystallise it from EtOH. [Beilstein 5 IV 738.]

1,3-Dinitrobenzene [99-65-0] M 168.1, m 90.5-91°. Crystallise 1,3-dinitrobenzene from alkaline EtOH solution (20g in 750ml 95% EtOH at 40°, plus 100ml of 2M NaOH) by cooling and adding 2.5L of  $H_2O$ . The precipitate, after filtering off, is washed with  $H_2O$ , sucked dry, and crystallised from 120ml, then 80ml of absolute EtOH [Callow et al. *Biochem J* 32 1312 1938]. Alternatively, crystallise it from MeOH,  $CCl_4$  or EtOAc. It can be sublimed in a vacuum. [Tanner *J Org Chem* 52 2142 1987, Beilstein 5 IV 739.]

1,4-Dinitrobenzene [100-25-4] M 168.1, m 173°. Crystallise 1,4-dinitrobenzene from EtOH or EtOAc. Dry it under vacuum over  $P_2O_5$ . It can be sublimed in a vacuum. [Beilstein 5 IV 741.]

2,4-Dinitrobenzenesulfonyl chloride [528-76-7] M 234.6, m 96°. Crystallise the sulfonyl chloride from  $CCl_4$ . [Beilstein 6 II 316.]

2,4-Dinitrobenzenesulfonyl chloride [1656-44-6] M 266.6, m 102°. Crystallise the sulfonyl chloride from \*benzene or \*benzene/petroleum ether. [Beilstein 11 H 78, 11 IV 214.]

2,4-Dinitrobenzoic acid [610-30-3] M 212.1, m 183°,  $pK^{25}$  1.42. Crystallise the acid from aqueous 20% EtOH (10ml/g) and let it dry at 100°. [Beilstein 9 II 279, 9 III 1776, 9 IV 1239.]

2,5-Dinitrobenzoic acid [610-28-6] M 212.1, m 179.5-180°,  $pK^{25}$  1.62. Crystallise the acid from distilled  $H_2O$ . Dry it in a vacuum desiccator. [Beilstein 9 II 279, 9 III 1778, 9 IV 1241.]

2,6-Dinitrobenzoic acid [603-12-3] M 212.1, m 202-203°,  $pK^{25}$  1.14. Crystallise the acid from water. [Beilstein 9 II 279, 9 III 1778, 9 IV 1242.]

3,4-Dinitrobenzoic acid [528-43-0] M 212.1, m 166°,  $pK^{25}$  2.81. Crystallise the acid from EtOH by addition of water. [Beilstein 9 III 1778, 9 IV 1242.]

**3,5-Dinitrobenzoic acid** [99-34-3] **M 212.1, m 205°, pK<sup>25</sup> 2.73 (2.79)**. Crystallise the acid from distilled H<sub>2</sub>O or 50% EtOH (4ml/g). Dry it in a vacuum desiccator or at 70° over BaO under a vacuum for 6 hours. [Beilstein 9 II 279, 9 III 1779, 9 IV 1242.]

**4,4'-Dinitrobenzoic anhydride** [902-47-6] **M 406.2, m 192°, 195-195.5°**. Crystallise the anhydride from Me<sub>2</sub>CO, toluene, \*C<sub>6</sub>H<sub>6</sub>/EtOAc or EtOAc. [Beilstein 9 H 393, 9 II 268, 9 III 1684.]

**3,5-Dinitrobenzoyl chloride** [99-33-2] **M 230.6, m 69.5°**. Crystallise it from CCl<sub>4</sub> or petroleum ether (b 40-60°). It reacts readily with H<sub>2</sub>O and should be kept in sealed ampoules. [Beilstein 9 IV 1350.]

**2,2'-Dinitrobiphenyl** [2436-96-6] **M 244.2, m 123-124°, b 194°/4mm**. Crystallise the biphenyl from EtOH, AcOH (m 124.5°) or petroleum ether (m 125.5-126°). [Beilstein 5 H 538, 5 III 1759, 5 IV 1826.]

**2,4'-Dinitrobiphenyl** [606-81-5] **M 244.2, m 92.7-93.7°**. Crystallise the biphenyl from EtOH. [Beilstein 5 H 538, 5 I 274, 5 II 491, 5 III 1759, 5 IV 1827.]

**4,4'-Dinitrobiphenyl** [1528-74-1] **M 244.2, m 240.9-241.8°**. Crystallise the biphenyl from \*C<sub>6</sub>H<sub>6</sub>, EtOH (charcoal) or Me<sub>2</sub>CO. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. It sublimes at ~138°/2.9×10<sup>-3</sup>mm. [Beilstein 5 H 584, 5 III 1760, 5 IV 1827.]

**2,6-Dinitro-*p*-cresol (2,6-dinitro-4-methylphenol)** [609-93-8] **M 198.1, m 78-79°, 80.5°, 81-82°, 85°, pK<sub>Est</sub> ~3.7**. Recrystallise the cresol from EtOH. It is steam volatile. The *piperidine salt* has m 195°. [Beilstein 6 H 414, 6 II 391, 6 III 1390, 6 IV 2152.] **TOXIC IRRITANT.**

**4,6-Dinitro-*o*-cresol (4,6-dinitro-2-methylphenol)** [534-52-1] **M 198.1, m 85-86°, 86-87°, 87°, pK<sup>25</sup> 4.70**. The cresol crystallises from aqueous EtOH. [Beilstein 6 H 369, 6 III 1276.]

**2,4-Dinitrodiphenylamine** [961-68-2] **M 259.2, m 157°, 160°, pK<sub>Est</sub> <0**. The amine forms red crystals from aqueous EtOH or CHCl<sub>3</sub>/EtOH (m 158°). The UV has λ<sub>max</sub> at 335nm (cyclohexane). [Beilstein 12 H 751, 12 III 1683, 12 IV 1692.]

**4,4'-Dinitrodiphenylurea [1,3-bis-(4-nitrophenyl)urea]** [587-90-6] **M 302.2, m 312°(dec)**. Crystallise the urea from EtOH (m 364°, long heating), EtOH/Me<sub>2</sub>CO (m 301-303°, 300-304° dec, 318-319°) or Me<sub>2</sub>CO (m 289° dec). It sublimes *in vacuo*. [Beilstein 12 H 723, 12 II 393, 12 III 1619, 12 IV 1646.]

**2,4-Dinitrofluorobenzene (Sanger's reagent)** [70-34-8] **M 186.1, m 25-27°, b 133°/2mm, 140-141°/5mm, d<sub>4</sub><sup>20</sup> 1.483**. Crystallise the reagent from Et<sub>2</sub>O or EtOH. Distil it in a vacuum through a Todd Column. If it is to be purified by distillation *in vacuo*, the distillation unit must be allowed to cool before air is allowed into the apparatus; otherwise the residue carbonises spontaneously and an **EXPLOSION** may occur. The material is a **skin irritant** and may cause serious dermatitis. [Beilstein 5 IV 742.]

**1,8-Dinitronaphthalene** [602-38-0] **M 218.2, m 170-171°, 171.3°**. Crystallise it from \*benzene. [Beilstein 5 H 559, 5 II 455, 5 III 1608.]

**2,4-Dinitro-1-naphthol (Martius Yellow)** [605-69-6] **M 234.2, m 81-82°, pK<sub>Est</sub> ~3.7**. Crystallise the naphthol from \*benzene or aqueous EtOH. [Beilstein 6 IV 4240.]

**2,4-Dinitrophenetole** [610-54-8] **M 240.2, m 85-86°, 210-211°/15mm**. Crystallise it from aqueous EtOH. The 1:1 *naphthalene complex* has m 41° and is obtained by fusing the compound with naphthalene in various ratios, then crystallising the solidified mix from a little EtOH (Dermer & Smith *J Am Chem Soc* 61 748 1939). [Beilstein 6 H 254, 6 III 858, 6 IV 1373.]

**2,4-Dinitrophenol** [51-28-5] **M 184.1, m 114°, pK<sup>25</sup> 4.12**. Crystallise it from \*benzene, EtOH, EtOH/H<sub>2</sub>O or H<sub>2</sub>O acidified with dilute HCl, dry it, then recrystallise it from CCl<sub>4</sub>. Dry it in an oven and store it in a vacuum

desiccator over CaSO<sub>4</sub>. The *benzoate* has **m** 132° (from EtOH). [*Beilstein* 6 IV 1369.]

**2,5-Dinitrophenol** [329-71-5] **M** 184.1, **m** 108°, **pK**<sup>25</sup> 5.20. Crystallise 2,5-dinitrophenol from H<sub>2</sub>O with a little EtOH. [*Beilstein* 6 IV 1383.]

**2,6-Dinitrophenol** [573-56-8] **M** 184.1, **m** 63.0-63.7°, **pK**<sup>25</sup> 3.73. Crystallise it from H<sub>2</sub>O, aqueous EtOH, \*C<sub>6</sub>H<sub>6</sub>/cyclohexane, or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 60-80°, 1:1). [*Beilstein* 6 III 867, 6 IV 1383.]

**3,4-Dinitrophenol** [577-71-9] **M** 184.1, **m** 138°, **pK**<sup>25</sup> 5.42. Steam distil and crystallise it from H<sub>2</sub>O then dry it in air. **EXPLOSIVE** when dry, store it with 10% H<sub>2</sub>O. [*Beilstein* 6 III 868, 6 IV 1384.]

**3,5-Dinitrophenol** [586-11-8] **M** 184.1, **m** 126°, **pK**<sup>25</sup> 6.68. Crystallise it from \*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub>/petroleum ether. Store it with 10% water as it is **EXPLOSIVE** when dry. [*Beilstein* 6 III 869, 6 IV 1383.]

**2,4-Dinitrophenylacetic acid** [643-43-6] **M** 226.2, **m** 179°(dec), **pK**<sup>25</sup> 3.50. Crystallise the acid from H<sub>2</sub>O. [*Beilstein* 9 IV 1691.]

**2,4-Dinitrophenylhydrazine (DNPH)** [119-26-6] **M** 198.1, **m** 200°(dec), **pK**<sub>Est</sub> ~2.0. Crystallise DNPH from butan-1-ol, dioxane, EtOH, \*C<sub>6</sub>H<sub>6</sub> or EtOAc. The *hydrochloride* has **m** 186° (dec). [*Beilstein* 15 IV 380.]

**2,4-Dinitroresorcinol** [519-44-8] **M** 200.1, **m** 212.5-214.5°, **pK**<sup>25</sup> 3.05 (3.79). Crystallise the resorcinol from aqueous EtOH. Its solubility at 25° is 0.25% in EtOH and 0.08g/L in H<sub>2</sub>O. **EXPLOSIVE**. [*Beilstein* 6 II 824, 6 III 4352, 6 IV 5696.]

**3,5-Dinitrosalicylic acid** [609-99-4] **M** 228.1, **m** 173-174°, **pK**<sub>1</sub><sup>25</sup> 0.70, **pK**<sub>2</sub><sup>25</sup> 7.40. Crystallise the acid from H<sub>2</sub>O. [*Beilstein* 10 IV 270.]

**2,6-Dinitrothymol** [303-21-9] **M** 240.2, **m** 54.5-55°, 80-81°, 81°. Crystallise 2,4-dinitrothymol from aqueous EtOH or petroleum ether. [Ganguly & LeFèvre *J Chem Soc* 851 1934, *Beilstein* 6 H 543, 6 III 1911.]

**2,3-Dinitrotoluene** [602-01-7] **M** 182.1, **m** 59.2°, 63°. Distil the toluene in steam and crystallise it from H<sub>2</sub>O or \*benzene/petroleum ether. Store it with 10% H<sub>2</sub>O as it could be **EXPLOSIVE** when dry. [*Beilstein* 5 H 339, 5 III 758, 5 IV 865.]

**2,4-Dinitrotoluene** [121-14-2] **M** 182.1, **m** 70.5-71.0°, 71.8-72.2°. Crystallise it from Me<sub>2</sub>CO, isopropanol or MeOH. Dry it in a vacuum over H<sub>2</sub>SO<sub>4</sub>. It has also been purified by zone melting. Store it with 10% of H<sub>2</sub>O. *It could be* **EXPLOSIVE** when dry. [*Beilstein* 5 H339, 5 IV 865, 5 III 759.]

**2,5-Dinitrotoluene** [619-15-8] **M** 182.1, **m** 50.3°, 51.2°. Crystallise it from \*benzene. Store it with 10% of H<sub>2</sub>O. **EXPLOSIVE** when dry. [*Beilstein* 5 III 760, 5 IV 866.]

**2,6-Dinitrotoluene** [606-20-2] **M** 182.1, **m** 64.3°, 66.1°. Crystallise it from acetone. Store it with 10% of H<sub>2</sub>O. **EXPLOSIVE** when dry. [*Beilstein* 5 III 761, 5 IV 866.]

**3,4-Dinitrotoluene** [610-39-9] **M** 182.1, **m** 58.5°, 60°, 61°. Steam distil it and crystallise it from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. Store it with 10% of H<sub>2</sub>O to avoid **EXPLOSION**. [*Beilstein* 5 H 341, 5 III 761, 5 IV 866.]

**3,5-Dinitro-*o*-toluic acid** [28169-46-2] **M** 226.2, **m** 206°, 207.5-208°, **pK**<sub>Est</sub> ~3.0. Crystallise the acid from H<sub>2</sub>O or aqueous EtOH. The *ammonium salt* forms yellow crystals from EtOH with **m** 218-219°, and the *urea salt* has **m** 189-190° (prisms from EtOH). [*Beilstein* 9 H 474, 9 II 323, 9 III 2316.]

**2,4-Dinitro-*m*-xylene** [603-02-1] **M** 196.2, **m** 83-84°. Crystallise the yellow 2,4-dinitro-*m*-xylene from EtOH. [*Beilstein* 5 H 379, 5 II 295, 5 III 844.]

**Dinonyl phthalate (mainly 3,5,5-trimethylhexyl phthalate isomer)** [14103-61-8, 28553-12-0, 84-76-4] **M 418.6, m 26-29°, b 170°/2mm,  $d_4^{20}$  0.9640,  $n_D^{20}$  1.4825.** Wash the ester with aqueous  $\text{Na}_2\text{CO}_3$  then shake it with water. Ether is added to break the emulsion, and the solution is washed twice with water, and dried ( $\text{CaCl}_2$ ). After evaporating the ether, the residual liquid is distilled three times under reduced pressure. It is stored in a vacuum desiccator over  $\text{P}_2\text{O}_5$ . [Beilstein 9 IV 3183.]

**4,4'-Di-*n*-pentyloxyazoxybenzene** [64242-26-8] **M 370.5, m 124.5° (with phase change at 82°, becomes clear and remelts at 119°).** Crystallise it from  $\text{Me}_2\text{CO}$ , and dry it by heating under vacuum. [Beilstein 16 III 599, 16 IV 175.]

**Diphenic acid (diphenyl-2,2'-dicarboxylic acid)** [482-05-3] **M 242.2, m 228-229°,  $\text{pK}^{25}$  3.46.** Crystallise diphenic acid from water. Hot  $\text{Ac}_2\text{O}$  provides the *anhydride* below. [Beilstein 9 H 922, 9 IV 3552.]

**Diphenic anhydride (diphenyl-2,2'-dicarboxylic anhydride)** [6050-13-1] **M 466.3, m 217°, 222-224°.** After removing free acid by extraction with cold aqueous  $\text{Na}_2\text{CO}_3$ , the residue is crystallised from acetic anhydride and dried at 100°. Acetic anhydride converts the acid to the anhydride. It also crystallises from  $^*\text{C}_6\text{H}_6$  (m 219°) or chlorobenzene (m 224.5-225.5°). [Beilstein 17 H 526, 17 II 495, 17 III/IV 6425.]

***N,N*-Diphenylacetamide** [621-09-0] **M 210.3, m 131°.** Crystallise it from EtOH, then sublime it under vacuum at *ca* 96° onto a "finger" cooled in solid  $\text{CO}_2/\text{MeOH}$ , with continuous pumping to free it from occluded solvent. [Beilstein 12 H 248, 12 II 144, 12 III 471, 12 IV 384.]

**Diphenylacetic acid** [117-34-0] **M 212.3, m 147.4-148.4°,  $\text{pK}^{25}$  3.94.** Crystallise the acid from  $^*$ benzene,  $\text{H}_2\text{O}$  or aqueous 50% EtOH. [Beilstein 9 H 673, 9 IV 2492.]

**Diphenylacetonitrile** [86-29-3] **M 193.3, m 73-75°.** Crystallise the nitrile from EtOH or petroleum ether (b 90-100°). [Beilstein 9 H 674, 9 IV 2505.]

**Diphenylacetylene (tolan)** [501-65-5] **M 178.2, m 62.5°, b 90-97°/0.3mm.** Crystallise tolan from EtOH. [Beilstein 5 H 656, 5 IV 2276.]

**Diphenylamine** [122-39-4] **M 169.2, m 62.0-62.5°,  $\text{pK}^{25}$  0.77 (aqueous  $\text{H}_2\text{SO}_4$ ).** Crystallise diphenylamine from petroleum ether, MeOH, or EtOH/water. Dry it under vacuum. [Beilstein 12 H 174, 12 IV 271.]

**Diphenylamine-2-carboxylic acid (*N*-phenylanthranilic acid)** [91-40-7] **M 213.2, m 182-183°, 184°,  $\text{pK}_1^{25}$  -1.28 (aqueous  $\text{H}_2\text{SO}_4$ ),  $\text{pK}_2^{25}$  3.86( $\text{CO}_2\text{H}$ ).** Crystallise the acid from EtOH (5ml/g) or AcOH (2ml/g) by adding hot water (1ml/g). [Beilstein 14 IV 1019.]

**Diphenylamine-2,2'-dicarboxylic acid (2,2'-iminodibenzoic acid)** [579-92-0] **M 257.2, m 298°(dec), 302°(dec), 320°(dec, long heating),  $\text{pK}_1^{35}$  6.05,  $\text{pK}_2^{35}$  7.02 (in 50% aqueous dioxane).** Crystallise the acid from EtOH. It complexes with  $\text{Cu}^{2+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Cd}^{2+}$ ,  $\text{Co}^{2+}$  and  $\text{Ni}^{2+}$ . [Beilstein 14 H 354, 14 I 545, 14 III 942, 14 IV 1058.]

**9,10-Diphenylanthracene** [1499-10-1] **M 330.4, m 248-249°.** Crystallise the anthracene from acetic acid or xylene [Baumstark et al. *J Org Chem* 52 3308 1987]. [Beilstein 5 IV 2807.]

***N,N'*-Diphenylbenzidine** [531-91-9] **M 336.4, m 245-247°, 251-252°,  $\text{pK}^{25}$  0.30.** Crystallise the benzidine from toluene or ethyl acetate. Store it in the dark. [Beilstein 13 H 223, 13 IV 368.]

***trans-trans*-1,4-Diphenylbuta-1,3-diene** [538-81-8] **M 206.3, m 153-153.5°.** Its solution in petroleum ether (b 60-70°) is chromatographed on an alumina-Celite column (4:1), and the column is washed with the same solvent. The main zone is cut out, eluted with ethanol and transferred to petroleum ether, which is then dried and evaporated [Pinckard et al. *J Am Chem Soc* 70 1938 1948]. Recrystallise it from hexane. [Beilstein 5 H 676, 9 IV 2319.]

***sym*-(1,5)-Diphenylcarbazide** [140-22-7] M 242.3, m 172°. A common impurity is phenyl-semicarbazide which can be removed by chromatography: ~8g in H<sub>2</sub>O is placed on a column (polyamide 6 powder, Macherey-Nagel-GmbH-Germany, washed several times with MeOH), eluted with H<sub>2</sub>O/MeOH/AcOH (1:3:0.04) at 7-8 drops/second, then eluted with the same solvent mixture but diluted 5 fold with H<sub>2</sub>O. The purification is followed by UV light at 280nm. The effluent is evaporated to dryness *in vacuo* at ~28°. [chromatography, IR & UV: Willems et al. *Anal Chim Acta* **51** 544 1970]. Recrystallise it from EtOH by adding CCl<sub>4</sub> to induce crystallisation, or AcOH to give a white crystalline powder which turns pink in air. It is air and light sensitive and should be stored in the dark under N<sub>2</sub>. [*Beilstein* **15** H 292, **15** IV 182.]

**1,5-Diphenylcarbazone (phenyldiazinecarboxylic acid 2-phenylhydrazide)** [538-62-5] M 240.3, m 124-127°, 156-159°(dec). It crystallises from EtOH (*ca* 5ml/g), and dry the orange-red needles at 50°. A commercial sample, nominally *sym*-diphenylcarbazone (m 154-156°) was a mixture of diphenylcarbazide and diphenylcarbazone. The former was removed by dissolving 5g of the crude material in 75ml of warm EtOH, then adding 25g Na<sub>2</sub>CO<sub>3</sub> dissolved in 400ml of distilled water. The alkaline solution was cooled and extracted six times with 50ml portions of diethyl ether (discarded). Diphenylcarbazone was then precipitated by acidifying the alkaline solution with 3M HNO<sub>3</sub> or glacial acetic acid. It was filtered off, air dried, and stored in the dark [Gerlach & Frazier *Anal Chem* **30** 1142 1958]. Other impurities are phenylsemicarbazide and diphenylcarbodiimide. Impurities can be detected by chromatography [Willems et al. *Anal Chim Acta* **51** 544 1970]. It is used for detection and estimation of Hg, Zn, Cd, Cr, Cu, Fe and Mo [Cheng et al. *Handbook of Organic Analytical Reagents*, Boca Baton 277 1982, *Beilstein* **16** H 24, **16** IV 17.]

**Diphenyl carbonate (phenyl carbonate)** [102-09-0] M 214.2, m 80°, 168°/15mm, 306°/760mm. Purify it by distillation under reduced pressure, sublimation, or by gas chromatography with 20% Apiezon on Embacel, and crystallisation from EtOH. [*Beilstein* **6** H 158, **6** IV 629.]

**Diphenylcyclopropenone (Diphencyprone)** [886-38-4] M 206.2, m 87-90°(hydrate), 119-121°(anhydrous). Crystallise it from cyclohexane. Its UV (MeCN) has  $\lambda_{\max}$  at 226, 282, 297nm. [*Beilstein* **17** IV 1736.]

**Diphenyl disulfide (phenyl disulfide)** [882-33-7] M 218.3, m 60.5°. Crystallise the disulfide from MeOH. [Alberti et al. *J Am Chem Soc* **108** 3024 1986]. Also crystallise it repeatedly from hot Et<sub>2</sub>O, then dry it in a vacuum at 30° over P<sub>2</sub>O<sub>5</sub>, fuse it under N<sub>2</sub> and re-dry it; the whole procedure being repeated, with a final drying under a vacuum for 24 hours. *Alternatively*, recrystallise it from hexane/EtOH solution. [Burkey & Griller *J Am Chem Soc* **107** 246 1985, *Beilstein* **6** H 323, **6** IV 1560.]

**1,1-Diphenylethanol** [599-67-7] M 198.3, m 80-81°, 81-81.5°, 87°, 90°, b 144-145°/12mm, 260°/760mm (dec),  $d_4^{15}$  1.1057. Crystallise 1,1-diphenylethanol from *n*-heptane and/or distil it under vacuum. The *benzoyl* derivative has m 115° and the *phenylurethane* has m 119°. [Bromberg et al. *J Am Chem Soc* **107** 83 1985, *Beilstein* **6** H 685, **6** I 330, **6** II 639, **6** III 3395, **6** IV 4713.]

**1,1-Diphenylethylene** [530-48-3] M 180.3, m 6°, b 101°/2mm, 134°/10mm, 268-270°/760mm,  $d_4^{20}$  1.024,  $n_D^{20}$  1.6088. Distil it under reduced pressure from KOH. Dry it with CaH<sub>2</sub> and redistil it. [*Beilstein* **5** H 639, **5** III 1975, **5** IV 2173.]

***N,N'*-Diphenylethylenediamine (Wanzlick's Reagent)** [150-61-8] M 212.3, m 67.5°, b 178-182°/2mm  $pK_{\text{Est}(1)} \sim 0.5$ ,  $pK_{\text{Est}(2)} \sim 3.8$ . Crystallise the reagent from aqueous EtOH or MeOH. [*Beilstein* **12** H 543, **12** IV 986.]

**(1*S*,2*S*)-1,2-diphenylethylenediamine bis-triflamide** [121788-77-0] M 478.4, m 213-214°,  $[\alpha]_D^{23} -6.6^\circ$  (c 1.4, CHCl<sub>3</sub>). The bis-amide is obtained from (1*S*,2*S*)-(-)-1,2-diphenylethylenediamine (see [*cis*-1-aminoindan-2-ol -69-8]) by treatment with triflic anhydride (358-23-6), Et<sub>3</sub>N and 4-dimethylaminopyridine (DMAP, see [1122-58-3]) in CH<sub>2</sub>Cl<sub>2</sub>; and purified by flash chromatography on silica gel (15% EtOAc-hexane, v/v). Its <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>, TMS) has  $\delta_H$  at 4.81(s, 2H), 6.80 (br s, 2H), 7.25 (s, 6H), 7.0 (s, 4H); and its <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta$  at 63.7, 127.0, 129.1 (2C), 135.1. It is a useful reagent for enantioselective Diels-Alder and aldol reactions using a chiral controller system. [Corey et al. *J Am Chem Soc* **111** 5493 1989, Pikul &

Corey *OrgSynth* 71 30 1993 and references herein.]

*N,N'*-Diphenylformamidine [622-15-1] M 196.2, m 142°, 137°, 136-139°. Crystallise it from absolute EtOH. The *hydrate* is obtained from aqueous EtOH. [Beilstein 12 H 236, 12 IV 372.]

1,3-Diphenylguanidine [102-06-7] M 211.3, m 148°, pK<sup>25</sup> 10.12. Crystallise it from toluene, aqueous acetone or EtOH, and dry it in a vacuum. [Beilstein 12 H 369, 12 IV 769.]

1,6-Diphenyl-1,3,5-hexatriene [1720-32-7] M 232.3, m 200-203°. Crystallise the hexatriene from CHCl<sub>3</sub> or EtOH/CHCl<sub>3</sub> (1:1). [Beilstein 5 H 691, 5 IV 2425.]

1,1-Diphenylhydrazine [530-50-7] M 184.2, m 34.5°, 44°, b 120°/1.1mm, 220°/40-50mm, pK<sub>Est</sub> ~9.1. Purify it via the *hydrochloride* [530-47-2], which has m 165-170°(dec) after crystallisation from aqueous EtOH (plus a few drops of HCl) and recover the free base with aqueous NaOH, extract it into Et<sub>2</sub>O, dry it (KOH), filter, evaporate and distil the residue under a vacuum. The distillate crystallises on cooling. The *benzoyl* derivative has m 194° (from Me<sub>2</sub>CO), and the *4-nitrophenyl hydrazone* (with the aldehyde) has m 131-132°. [Koga & Anselme *J Org Chem* 33 3963 1968, Beilstein 15 IV 55.]

1,2-Diphenylhydrazine (hydrazobenzene) [122-66-7] M 184.2, m 124-126°, b 175°/10mm, 222°/40mm, pK<sub>Est</sub> ~1.7. Crystallise hydrazobenzene from hot EtOH containing a little ammonium sulfide or H<sub>2</sub>SO<sub>3</sub> (to prevent atmospheric oxidation), preferably under N<sub>2</sub>. Dry it in a vacuum desiccator, and store it in the dark or under N<sub>2</sub>. It has been distilled in a vacuum. *Alternatively*, crystallise it from petroleum ether (b 60-100°) to constant absorption spectrum. It is almost colourless but in air it turns yellow, then red with the formation of azobenzene. The *hydrochloride* crystallises from EtOH and has m 163-164°(dec); however, hydrazobenzene readily rearranges to *benzidine* in the presence of acid. The *picrate* crystallises from \*C<sub>6</sub>H<sub>6</sub> and has m 123°(dec). [Beilstein 15 H 123.]

Diphenylmethane [101-81-5] M 168.2, m 25.4°, b 74°/0.4mm, 260.5°/764.5mm. Sublime it under vacuum, or distil it at 72-75°/0.4mm. Recrystallise it from cold EtOH. It has also been purified by fractional crystallisation from the melt. [Armarego *Aust J Chem* 13 95 1960, Beilstein 5 II 498, 5 IV 1841.]

1,1-Diphenylmethylamine (benzhydrylamine, aminodiphenylmethane) [91-00-9] M 183.2, m 12°, 34°, b 166°/12mm, 295°/atm, d<sub>4</sub><sup>25</sup> 1.063, n<sub>D</sub><sup>20</sup> 1.596, pK<sub>Est</sub> ~9.1. Crystallise the amine from H<sub>2</sub>O. The *free base* absorbs CO<sub>2</sub> from the atmosphere; store it accordingly. The *hydrochloride* [5267-34-5] M 219.7 m 293-295°, crystallises from H<sub>2</sub>O. [Beilstein 12 H 1323, 12 IV 3282.]

§A polymer bound *benzhydrylamine hydrochloride* is commercially available.

Diphenylmethyl chloride (benzhydryl chloride) [90-99-3] M 202.7, m 17.0°, b 140°/3mm, 167°/17mm, n 1.5960. Dry the chloride with Na<sub>2</sub>SO<sub>4</sub> and fractionally distil it under reduced pressure. [Beilstein 5 H 590, 5 I 278, 5 II 500, 5 III 1790, 5 IV 1847.]

*all-trans*-1,8-Diphenyl-1,3,5,8-octatetraene [3029-40-1] M 258.4, m 235-237°. Crystallise the octatetraene from EtOH. [Beilstein 5 H 709, 5 II 620, 5 III 2328, 5 IV 2508.]

*N,N'*-Diphenyl-*p*-phenylenediamine [74-31-7] M 260.3, m 148-149°, b 219-224°/0.7mm, pK<sub>Est</sub> <0. Crystallise the diamine from EtOH, chlorobenzene/petroleum ether or \*benzene. It has also been crystallised from aniline, then extracted three times with absolute EtOH. [Beilstein 13 H 80.]

1,1-Diphenyl-2-picrylhydrazine [1707-75-1] M 395.3, m 174°(dec), 178-179.5°(dec). Crystallise the hydrazine from CHCl<sub>3</sub>, or \*benzene/petroleum ether (1:1), then degas it at 100° and <10<sup>-5</sup>mm Hg for *ca* 50 hours to decompose the 1:1 molar complex formed with \*benzene. [Beilstein 15 II 221, 15 IV 1210.]

2,2-Diphenylpropionic acid [5558-66-7] M 226.3, m 173-174°, pK<sub>Est</sub> ~3.8. Crystallise the acid from EtOH. [Beilstein 9 II 474.]

**3,3-Diphenylpropionic acid** [606-83-7] M 226.3, m 155°, pK<sub>Est</sub> ~4.5. Crystallise the acid from EtOH. [Beilstein 9 H 680.]

**Diphenyl sulfide** [139-66-2] M 186.3, b 145°/8mm, d<sub>4</sub><sup>20</sup> 1.114, n<sub>D</sub><sup>20</sup> 1.633. Wash the sulfide with aqueous 5% NaOH, then water. Dry it with CaCl<sub>2</sub>, then with sodium. The sodium is filtered off, and the diphenyl sulfide is distilled under reduced pressure. [Beilstein 2 H 299, 6 IV 1488.]

**Diphenyl sulfone** [127-63-9] M 218.3, m 125°, b 378°(dec). Crystallise the sulfone from diethyl ether. It has been purified by zone melting. [Beilstein 6 H 300, 6 IV 1490.]

**sym-Diphenylthiourea (thiocarbanilide)** [102-08-9] M 228.3, m 154°. Crystallise the thiourea from boiling EtOH by adding hot water and allowing to cool. [Beilstein 12 H 394, 12 IV 810.]

**1,1-Diphenylurea** [603-54-3] M 212.3, m 238-239°. Crystallise 1,1-diphenylurea from MeOH. [Beilstein for 1,3 12 IV 741.]

**2,6-Di-*iso*-propylaniline** [24544-04-5] M 177.3, m -45°, b 120-122°/10mm, 257°/atm, d<sub>4</sub><sup>25</sup> 0.9367, n<sub>D</sub><sup>20</sup> 1.5330, pK<sub>Est</sub> ~0. Nitration of 1,3-di-*iso*-propylbenzene in AcOH/Ac<sub>2</sub>O (2:1 v/v) with 96% HNO<sub>3</sub> (1.24-2.05 equivalents) below 45-50°/24 hours gave, on fractionation through a 2.5 x 180cm column at 5-10mm, a mixture of lower boiling 4-nitro- (74%) and the higher boiling 2-nitro-1,3-di-*iso*-propylbenzene (24%). Reduction of the later in *iso*-propanol with H<sub>2</sub> (~1200psi, <100°) catalysed by Raney Ni (5-10% by weight), followed by filtration, evaporation of the solvent and fractionation through a 180cm column at 10mm, with a reflux ratio of 20:1 and a rate of 0.33ml/minute gave 2,6-di-*iso*-propylaniline as a colourless oil in 94.9% yield. The *benzoyl derivative* separated as an oil which solidifies and provides a white crystalline powder, m 106.0-106.7°, on recrystallisation from *iso*-octane. [Newton *J Am Chem Soc* 65 2434 1943.] Alternatively, aniline (300g), AlCl<sub>3</sub> (18g), and Al powder (6g) are heated in a pressure vessel at 290°, and propylene is pumped into the vessel to a pressure of 250 atmospheres. From time to time more propylene is added to sustain the pressure. After ca 8 hours the pressure ceased to decrease, the vessel is opened and the product is shaken with dilute aqueous NaOH and fractionated under a vacuum. The three fractions that separate are 2,6-di-*iso*-propylaniline (78%), 2-*iso*-propylaniline (15%), and a small amount of 2,4,6-tri-*iso*-propylaniline [Stroh et al. *Angew Chem* 69 124 1957, also report m 184.5-187° for the *acetyl derivative*, and m 254-256° for the *benzoyl derivative*.] [Beilstein 12 III 2764.]

It is used for making NHCs for metal catalysed  $\alpha$ -arylation of acyclic ketones, e.g. propiophenones, and amination of haloarenes [Matsubara et al. *J Org Chem* 72 5069 2007]. It has been sulfonated for making sulfonated anilines, and consequently sulfonated NHCs, which provide water soluble Pd-NHCs used for Suzuki coupling of arylhalides with arylboronic acids in aqueous medium [Fleckenstein et al. *Chem Commun* 2870 2007].

**N-Trimethylsilyl 2,6-di-*iso*propylaniline** [78923-65-6], prepared by reaction of Me<sub>3</sub>SiCl and the Li salt of the aniline in Et<sub>2</sub>O, is purified by distillation in vacuum. It is used in preparing the key intermediate for the Schrock-Hoveyda molybdenum catalyst (see [205815-80-1]), and sometimes it is prepared *in situ* without isolating it while using Et<sub>3</sub>N instead of Li as base. [Schrock et al. *J Am Chem Soc* 112 3875 1990, Alexander et al. *Organometallics* 19 3700 2000].

**N,N-Di-*n*-propylaniline** [2217-07-4] M 177.3, b 127°/10mm, 238-241°/760mm, pK<sub>23</sub> 5.68. Reflux the aniline for 3 hours with acetic anhydride, then fractionally distil under reduced pressure.

**Dithizone (diphenylthiocarbazone)** [60-10-6] M 256.3, m 168°(dec), ratio of  $\epsilon_{620\text{nm}}/\epsilon_{450\text{nm}}$  should be  $\geq 1.65$ ,  $\epsilon_{620}$   $3.4 \times 10^4$  (CHCl<sub>3</sub>), pK<sub>2</sub> 4.6. The crude dithizone is dissolved in CCl<sub>4</sub> to give a concentrated solution. This is filtered through a sintered glass funnel and shaken with 0.8M aqueous ammonia to extract dithizonate ion. The aqueous layer is washed with several portions of CCl<sub>4</sub> to remove undesirable materials. The aqueous layer is acidified with dilute H<sub>2</sub>SO<sub>4</sub> to precipitate pure dithizone. This is dried in a vacuum. When only small amounts of dithizone are required, purification by paper chromatography is convenient. [Cooper & Hibbits *J Am Chem Soc* 75 5084 1933.] Instead of CCl<sub>4</sub>, CHCl<sub>3</sub> can be used, and the final extract, after washing with water, can be evaporated in air at 40-50° and dried in a desiccator. It complexes with Cd, Hg, Ni and Zn.

[*Beilstein* 16 H 26, 16 IV 18.]

**Di-*p*-tolyl carbonate** [621-02-3] **M 242.3, m 115°**. Purify the carbonate by GLC with 20% Apiezon on Embacel followed by sublimation *in vacuo* and recrystallisation from EtOH (**m 114°**). [*Beilstein* 6 H 398, 6 I 201, 6 II 380, 6 III 1366.]

***N,N'*-Di-*o*-tolylguanidine** [97-39-2] **M 239.3, m 179° (175-176°), pK<sup>18</sup> 9.62**. The guanidine crystallises from aqueous EtOH. The *sulfate* has **m 253-254° (dec, H<sub>2</sub>O)**. [*Beilstein* 12 H 803, 12 II 445, 12 III 1871, 16 IV 1764.]

**Di-*p*-tolylphenylamine** [20440-95-3] **M 273.4, m 108.5°, b 250°/20-40mm, pK<sub>Est</sub> ~ -5.0**. Crystallise the amine once from EtOAc, then twice from EtOH to give pale yellow needles. [Marsden *J Chem Soc* 627 1937, *Beilstein* 12 III 2033.]

**Di-*p*-tolyl sulfone** [599-66-6] **M 278.3, m 158-159°, 163.5°, b 405°/760mm**. Crystallise the sulfone repeatedly from Et<sub>2</sub>O or EtOH. It has been purified by zone melting. [*Beilstein* 6 H 419, 6 II 395, 6 III 1405, 6 IV 2174.]

**1,2-Divinylbenzene** [1321-74-0] **M 130.2, b 53°/3mm, 79.2-79.6°/15mm, 195°//760mm, d<sup>20</sup> 0.919, n<sub>D</sub><sup>20</sup> 1.573**. Purify divinylbenzene by dissolving in Et<sub>2</sub>O, shaking with H<sub>2</sub>O, drying over CaCl<sub>2</sub>, filtering, evaporating and distilling *in vacuo*. It polymerises within 2-3days unless 4-*tert*-butylcatechol (0.05%) is added as stabiliser. [Fries & Bestian *Chem Ber* 69 715 1936, *Beilstein* 5 III 1366.]

**Duroquinone (tetramethylbenzoquinone)** [527-17-3] **M 164.2, m 110-111°, 111-112°**. Crystallise duroquinone from 95% EtOH. Dry it *in vacuo*. [*Beilstein* 7 H 669, 7 III 3417, 7 IV 2101.]

**Ellagic acid (4,4',5,5',6,6'-hexahydroxydiphenic acid dilactone)** [476-66-4] **M 338.2, m >360°, pK<sub>Est(1)</sub> ~ 8, pK<sub>Est(2)</sub> ~ 11**. This antioxidant crystallises from pyridine. It forms a dark green solution in aqueous *N* NaOH. The *tetraacetate dilactone* crystallises from Ac<sub>2</sub>O, with **m 340°**. [*Beilstein* 19 H 261, 19 III/IV 3164, 19/7 V 108.]

**Emodine (1,3,8-trihydroxy-6-methyl-9,10-anthracenedione, archin)** [518-82-1] **M 270.2, m 253-257°, 255-256°, 256-257°, 262°, 264° (phenolic pKs 7—10)**. Archin forms orange needles from EtOH, Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub>, toluene or pyridine. It sublimes above 200° at 12mm. [Tutin & Clewer *J Chem Soc* 99 946 1911, IR: Bloom et al. *J Chem Soc* 178 1959, UV: Birkinshaw *Biochem J* 59 495 1955, Raistrick *Biochem J* 34 159 1940.]

**1*R*,2*S*-(-)Ephedrine** see (-)-ephedrine (1*R*,2*S*-2-methylamino-1-phenylpropanol) in "Miscellaneous" in Chapter 7.

**(±)-Epichlorohydrin** [106-89-8] **M 92.5, b 115.5°/760mm d<sub>4</sub><sup>25</sup> 1.183, n<sub>D</sub><sup>20</sup> 1.438**. Distil epichlorohydrin at 760mm, heat it on a steam bath with one-quarter its weight of CaO, then decant and fractionally distil it. [*Beilstein* 17 H 6, 17 III/IV 20.]

***R*-(-)Epinephrine** See adrenalin in "Miscellaneous" in Chapter 7.

**3-Ethoxy-*N,N*-diethylaniline** [1846-92-2] **M 193.3, b 145°/14mm, 268-270°/760mm, n<sub>D</sub><sup>25</sup> 1.5325, pK<sub>Est</sub> ~ 6.1.** Reflux it for 3 hours with Ac<sub>2</sub>O, then fractionally distil it, preferably under reduced pressure. The 3-*nitroso* derivative has **m 72.5°** (green crystals from petroleum ether/\*C<sub>6</sub>H<sub>6</sub>). [Klaassens & Schoot *Rec Trav Chim Pays Bas* 48 1272 1929.]

**1-Ethoxynaphthalene** [5328-01-8] **M 172.2, b 136-138°/14mm, 282°/760mm, d<sub>4</sub><sup>20</sup> 1.061, n<sub>D</sub><sup>20</sup> 1.604**. Fractionally distil it (twice) under a vacuum, then dry it with, and distil it under a vacuum from sodium. The *picrate* has **m 118.5-119°** (from EtOH). [*Beilstein* 6 H 606, 6 II 578, 6 III 2924, 6 IV 4212.]

**2-Ethoxynaphthalene** [93-18-5] **M 172.2, m 35.6-36.0°, b 142-143°/12mm, 280°/760mm.** Crystallise it from petroleum ether or EtOH (**m 37-38°**). Dry it *in vacuo*, or distil it in a vacuum. The *picrate* has **m 104.5°** (from EtOH or CHCl<sub>3</sub>). [Beilstein 6 H 606, 6 II 578, 6 III 2972, 6 IV 4257.]

**Ethyl *p*-aminobenzoate (Benzocaine)** [94-09-7] **M 165.2, m 92°, pK<sup>25</sup> 2.39.** Crystallise Benzocaine from EtOH/H<sub>2</sub>O or EtOH (**m 93-94°**), and dry it in air. [Beilstein 14 H 422, 14 IV 1129.]

***p*-Ethylaniline** [589-16-2] **M 121.2, b 88°/8mm, d<sub>4</sub><sup>20</sup> 0.975, n<sub>D</sub><sup>20</sup> 1.554, pK<sup>25</sup> 5.00.** Dissolve *p*-ethylaniline in \*benzene, then acetylate it. Recrystallise the *acetyl* derivative (**m 93-94°**, Hickinbottom & Waine *J Chem Soc* 1565 1930) from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or EOH, and hydrolyse it by refluxing 50g with 500ml of H<sub>2</sub>O and 115ml of conc H<sub>2</sub>SO<sub>4</sub> until the solution becomes clear. The amine sulfate is isolated, suspended in H<sub>2</sub>O and solid KOH is added to generate the free base, which separates. Dry it (KOH), and distil it from zinc dust in a vacuum [Berliner & Berliner *J Am Chem Soc* 76 6179 1954]. The *picrate* has **m 171-172°** (from 1,2-dichloroethane). [Beilstein 12 H 1090, 12 III 2380, 12 IV 2419.]

***trans*-Ethyl cinnamate** [103-36-6] **M 176.2, f 6.7°, b 127°/6mm, 272.7°/768mm, d<sub>4</sub><sup>20</sup> 1.040, n<sub>D</sub><sup>20</sup> 1.55983.** Wash the ester with aqueous 10% Na<sub>2</sub>CO<sub>3</sub>, then water, dry (MgSO<sub>4</sub>), and distil it. The purified ester is saponified with aqueous KOH, and, after acidifying the solution, cinnamic acid is isolated, washed and dried. The ester is reformed by refluxing for 15 hours the cinnamic acid (25g) with absolute EtOH (23g), conc H<sub>2</sub>SO<sub>4</sub> (4g) and dry \*benzene (100ml), after which it is isolated, washed, dried and distilled under reduced pressure [Jeffery & Vogel *J Chem Soc* 658 1958]. [Beilstein 9 IV 2006.]

**(±)-Ethyl α,β-dibromo-β-phenylpropionate** [5464-70-0, *erythro* 30983-70-1] **M 336.0, m 75°.** Crystallise the propionate from petroleum ether (b 60-80°), EtOH or aqueous EtOH (**m 78°**, and an unstable form **m 65°**). [Beilstein 9 H 512, 9 I 202, 9 III 2406, 9 IV 1772.]

**Ethylene di(*p*-toluenesulfonate) (ethylene glycol ditosylate)** [6315-52-2] **M 370.44, m 124-127°, m 126-127°, m 128°.** It is prepared from ethylene glycol (0.5mole) and *p*-toluenesulfonyl chloride (1.0mole) are stirred in Me<sub>2</sub>CO (100ml) in the presence of 25% aqueous NaOH (175ml, 44g, 1.1mole) at 2-10° for ~4 hours. Excess of ice-water is added, the solid is filtered off, washed with large volumes of H<sub>2</sub>O, dried *in vacuo* and recrystallised from dry \*C<sub>6</sub>H<sub>6</sub> (absolute EtOH has also been used) to give the *di-tosylate* (**m 128°**) in 76% yield [Drahowzal & Klamann *Monatsh* 82 452 1951]. Alternatively, ethylene glycol (0.25mole), *p*-toluenesulfonyl chloride (0.52mole) in dry pyridine (80g, 1mole) are stirred gently at -10° to -4° for 2 hours then treated with excess of ice-water as above gave the *diester* in 87.3% yield [Drahowzal & Klamann *Monatsh* 82 460 1951]. The latter method is superior as there is no H<sub>2</sub>O in the reaction medium. Note that attempts to prepare the *mono-p-tosylate ester* using equimolar amounts of glycol and sulfonyl chloride always produced the *diester* which crystallised as white plates (**m 126°**) from hot EtOH [Butler et al. *J Am Chem Soc* 57 575 1935]. [Beilstein 11 II 296, 11 III 225, 11 IV 290.]

**S,S'-Ethylene di(*p*-toluenethiosulfonate) (ethylene dithiotosylate)** [2225-23-2] **M 402.5, m 75-76°.** This reagent is the lower homologue of 1,3-trimethylene di(*thiotosylate*) [3866-79-3], and is made from potassium thiotosylate (45g, 0.20 mole, see below [28519-50-8]) and 1,2-dibromoethane (18.8g, 0.1 mole) in EtOH (200ml, containing 10-20mg of KI to activate the dibromide) by refluxing, with stirring in the dark under a N<sub>2</sub> atmosphere, for 8 hours. The solvent is evaporated *in vacuo*, and the white solid residue is washed, by decantation, with a mixture of EtOH (80ml) and H<sub>2</sub>O (150ml), and then with H<sub>2</sub>O (3 x 50ml), and recrystallised from EtOH (~150ml) to provide the crude thioester (28.7g) with m 72-75°. The pure thioester (24g, 60%), **m 75-76°**, is obtained as white crystals after three recrystallisations from EtOAc/EtOH; and its <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has δ at 2.47 (s, 6H, 2CH<sub>3</sub>), 3.31 (s, 4H, CH<sub>2</sub>CH<sub>2</sub>), 7.48 (d, *J* = 9Hz, 4H, Aromatic H) and 7.97 (d, *J* = 9Hz, 4H, Aromatic H). [Woodward et al. *Org Synth Coll Vol V* 1016 1988, *Org Synth* 54 33 1974.]

Like the trimethylene homologue below it can form 1,3-dithiolanes with activated methylene groups and carbonyl compounds.

**Ethyl gallate** [831-61-8] **M 198.2, m 150-151°, 163-165°.** Recrystallise the gallate from 1,2-dichloroethane. Its UV has λ<sub>max</sub> (neutral species) at 275nm (ε 10,000), (anion) at 235nm (ε 10,300), 279nm (ε 11,400) and

324nm ( $\epsilon$  8 500) [Campbell & Coppinger *J Am Chem Soc* **73** 2708 1951]. [Beilstein **10** IV 2002.]

**Ethyl *p*-nitrobenzoate** [99-77-4] **M 195.2, m 56°**. Dissolve it in Et<sub>2</sub>O and wash it with aqueous alkali, then the ether is evaporated and the solid recrystallised from EtOH. [Beilstein **13** H 3787, **13** IV 1074.]

***o*-Ethylphenol** [90-00-6] **M 122.2, f 45.1°, b 210-212°, d<sub>4</sub><sup>20</sup> 1.020, n<sub>D</sub><sup>20</sup> 1.537, pK<sup>25</sup> 10.20**. Purify as for *p*-ethylphenol below. [Beilstein **6** H 470, **6** IV 3011.]

***p*-Ethylphenol** [123-07-9] **M 122.2, m 47-48°, b 76°/1mm, 153°/100mm, 218.0°/762mm, n<sub>D</sub><sup>25</sup> 1.5239, pK<sup>25</sup> 10.21**. Non-acidic impurities are removed by passing steam through a boiling solution containing 1 mole of the phenol and 1.75 moles of NaOH (as an aqueous 10% solution). The residue is cooled and acidified with 30% (v/v) H<sub>2</sub>SO<sub>4</sub>, and the free phenol is extracted into diethyl ether. The extract is washed with water, dried with CaSO<sub>4</sub> and the ether is evaporated. The phenol is distilled at 100mm pressure through a Stedman gauze-packed column. It is further purified by fractional crystallisation by partial freezing, and by zone refining, under N<sub>2</sub> [Biddiscombe et al. *J Chem Soc* 5764 1963]. Alternatively, purify it via the benzoate, as for phenol. The 4-nitrophenylbenzoate has **m 80°** (from EtOH). [Beilstein **6** H 470, **6** III 1663, **5** IV 3020.]

**Ethyl phenylacetate** [101-97-3] **M 164.2, b 99-99.3°/14mm, d<sub>4</sub><sup>20</sup> 1.030, n<sub>D</sub><sup>20</sup> 1.499**. Shake the ester with saturated aqueous Na<sub>2</sub>CO<sub>3</sub> (three times), aqueous 50% CaCl<sub>2</sub> (twice) and saturated aqueous NaCl (twice). Dry with CaCl<sub>2</sub> and distil it under reduced pressure. [Beilstein **9** H 434, **9** IV 1618.]

**Ethyl Red [2-(4-diethylaminophenylazo)benzoic acid]** [76058-33-8] **M 297.4, m 135°, 150-152°, pK<sub>1</sub> 2.5, pK<sub>2</sub> 9.5**. Crystallise the acid dye from EtOH/diethyl ether or toluene. It is an indicator: pH 4.4 (red) and 6.2 (yellow). Its UV has  $\lambda_{\text{max}}$  at 447nm. [Beilstein **16** I 316.]

**Ethynyl *p*-tolylsulfone** [13894-21-8] **M 180.2, m 65-67°, 73-74°**. Recrystallise the sulfone from petroleum ether, \*C<sub>6</sub>H<sub>6</sub> or EtOH (m 66°), and dry it in a vacuum. [Beilstein **6** III 1397, **6** IV 2160.]

**Eugenol (4-allyl-2-methoxyphenol)** [97-53-0] **M 164.2, b 253°/760mm, 255°/760mm, d<sub>4</sub><sup>20</sup> 1.066, n<sub>D</sub><sup>20</sup> 1.540, pK<sup>25</sup> 10.19**. Fractional distillation of eugenol gives a pale yellow liquid which darkens and thickens on exposure to air. It should be stored under N<sub>2</sub> at -20°. [Waterman & Priedster *Rec Trav Chim Pays Bas* **48** 1272 1929, Beilstein **6** H 961, **6** IV 6337.]

**Eugenol methyl ether (4-allyl-1,2-dimethoxybenzene)** [93-15-2] **M 178.2, m -4°, b 127-129°/11mm, 146°/30mm, 154.7°/760mm, d<sub>4</sub><sup>20</sup> 1.0354, n<sub>D</sub><sup>20</sup> 1.53411**. Recrystallise the ether from hexane at low temperature and redistil it (preferably *in vacuo*). [Hillmer & Schorning *Z Phys Chem [A]* **167** 407 1934, Briner & Fliszár *Helv Chim Acta* **42** 2063 1959, Beilstein **6** H 963, **6** IV 6337.]

**Fluoranthene (benzo[*j,k*]fluorene)** [206-44-0] **M 202.3, m 110-111°, b 384°/760mm**. Purify it by chromatography of CCl<sub>4</sub> solutions on alumina, with \*benzene as eluent. Crystallise it from EtOH, MeOH or \*benzene. Also purify it by zone melting. [Gorman et al. *J Am Chem Soc* **107** 4404 1985, Beilstein **5** I 344, **5** IV 2463.]

**Fluorene** [86-73-7] **M 166.2, m 114.7-115.1°, b 160°/15mm, 298°/atm**. Purify fluorene by chromatography of CCl<sub>4</sub> or petroleum ether (b 40-60°) solution on alumina, with \*benzene as eluent. Crystallise it from 95% EtOH, 90% acetic acid and again from EtOH. Crystallisation using glacial acetic acid retains an impurity which is removed by partial mercuration and precipitation with LiBr [Brown et al. *J Am Chem Soc* **84** 1229 1962]. It has also been crystallised from hexane, or \*benzene/EtOH, distilled under vacuum and purified by zone refining. [Gorman et al. *J Am Chem Soc* **107** 4404 1985, Beilstein **5** IV 2142.]

**Fluorene-2,7-diamine** [525-64-4] **M 196.3, m 165-166°**. Crystallise the diamine from hot H<sub>2</sub>O or aqueous EtOH, dry it *in vacuo* and store it in the dark. [Beilstein **13** H 266, **13** II 123, **12** III 507, **13** IV 449.]

**9-Fluorenone** [486-25-9] **M 180.2, m 82.5-83.0°, 85-86°, b 341°/760mm.** Crystallise 9-fluorenone from absolute EtOH, MeOH or \*benzene/pentane. [Ikezawa *J Am Chem Soc* **108** 1589 1986.] Also recrystallise it twice from toluene and sublime it in a vacuum [Saltiel *J Am Chem Soc* **108** 2674 1986]. It can be distilled under high vacuum. The *oxime* has **m 195°, 198°** (yellow crystals from \*C<sub>6</sub>H<sub>6</sub> or xylene, Anet et al. *Can J Chem* **35** 180 1957, Wislicenus & Waldmuller *Chem Ber* **41** 3335 1908.) [Beilstein 7 H 465, 7 III 2330, 7 IV 1629.]

**9-Fluorenylmethyl chloroformate (Fmoc-Cl)** [28920-43-6] **M 258.7, m 61-63°, 61.4-63°.** If the IR contains no OH bands (at ~3000 cm<sup>-1</sup>) due to the hydrolysis product 9-fluorenylmethanol, then purify it by recrystallisation from dry Et<sub>2</sub>O. Its IR (CHCl<sub>3</sub>) has a band at 1770 cm<sup>-1</sup> (C=O), and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has δ at 4-4.6 (m 2H, CHCH<sub>2</sub>) and 7.1-7.8 (m, 8 aromatic H). The *azide* (Fmoc-N<sub>3</sub>) has **m 89-90°** (from hexane) and IR (CHCl<sub>3</sub>) at 2135 (N<sub>3</sub>) and 1730 (C=O) cm<sup>-1</sup>, and the *carbazate* (Fmoc-NHNH<sub>2</sub>) has **m 171°(dec)** (from nitromethane) with IR (KBr) at 3310, 3202 (NH) and 1686 (CONH) cm<sup>-1</sup>. [Caprino & Han *J Org Chem* **37**, 3404 1972, *J Am Chem Soc* **92** 5748 1970, Koole et al. *J Org Chem* **59** 1657 1989, Fürst et al. *J Chromatogr* **499** 537 1990.]

**9-Fluorenylmethyl succinimidyl carbonate** [82911-69-1] **M 337.3, m 147-151°(dec), 151° (dec).** Recrystallise the carbonate from CHCl<sub>3</sub>/Et<sub>2</sub>O, or from petroleum ether (b 40-60°). [Pauet *Can J Chem* **60** 976 1982, Lapatsaris et al. *Synthesis* 671 1983.]

**Fluorobenzene** [462-06-6] **M 96.1, b 84.8°/760mm, d<sub>4</sub><sup>20</sup> 1.025, n<sub>D</sub><sup>20</sup> 1.46573, n<sub>D</sub><sup>30</sup> 1.4610.** Dry fluorobenzene for several days with P<sub>2</sub>O<sub>5</sub>, then fractionally distil it. [Beilstein 5 H 198, 5 IV 632.]

***o*-Fluorobenzoic acid** [445-29-4] **M 140.1, m 127°, d<sub>25</sub> 1.460, pK<sub>25</sub> 3.27.** Crystallise the acid from 50% aqueous EtOH, dilute HCl or \*C<sub>6</sub>H<sub>6</sub>, then purify it by zone melting or vacuum sublimation at 130-140°. [Beilstein 9 H 333, 9 III 1324, 9 IV 950.]

***m*-Fluorobenzoic acid** [445-38-9] **M 140.1, m 124°, 125°, d<sub>25</sub> 1.474 pK<sub>25</sub> 3.86.** Crystallise the acid from 50% aqueous EtOH or \*C<sub>6</sub>H<sub>6</sub>, then sublime it *in vacuo* at 130-140°. [Beilstein 9 H 333, 9 IV 952.]

***p*-Fluorobenzoic acid** [456-22-4] **M 140.1, m 182°, d<sub>25</sub> 1.479, pK<sub>25</sub> 4.15.** Crystallise the acid from 50% aqueous EtOH, then purify it by zone melting or vacuum sublimation at 130-140°. [Beilstein 9 H 333, 9 III 1327, 9 IV 953.]

**3-Fluoro-4-hydroxyphenylacetic acid** [458-09-3] **M 170.1, m 33°, pK<sub>Est(1)</sub> ~4.4, pK<sub>Est(2)</sub> ~9.4.** Crystallise the acid from water. [Beilstein 10 III 440.]

**4-Fluoro-2-methylbenzaldehyde** [63082-45-1] **M 138.1, d<sub>4</sub><sup>25</sup> 1.144, n<sub>D</sub><sup>25</sup> 1.526.** The aldehyde has been purified by gas chromatography and should be kept under N<sub>2</sub> as it readily oxidises in air [Burgess et al. *Aust J Chem* **30** 543 1977].

**2-Fluoro-4-nitroaniline** [369-35-7] **M 156.1, m 134-135°, 135-136°, pK<sub>25</sub> -0.44 (aqueous HClO<sub>4</sub>).** The aniline forms yellow crystals on recrystallisation from aqueous MeOH or EtOH. The *acetyl* derivative has **m 203-204°** (from EtOH). [Wepster & Verkade *Rec Trav Chim Pays Bas* **68** 86 1949, Beilstein 12 III 1647.]

**1-Fluoro-4-nitrobenzene** [350-46-9] **M 141.1, m 27°(stable form), 21.5°(unstable form), b 86.6°/14mm, 95-97.5°/22mm, 205.3°/735mm.** Crystallise it from EtOH or distil it in a vacuum. [Beilstein 5 H 241, 5 IV 719.]

**1-Fluoro-4-nitronaphthalene** [341-92-4] **M 191.2, m 80°.** It crystallises from EtOH as yellow needles [Bunce et al. *J Org Chem* **52** 4214 1987]. [Beilstein 5 III 1596, 5 IV 1675.]

**4-Fluoro-3-nitrophenylazide** [28166-06-5] **M 182.1, m 53-55°, 54-56°.** Dissolve the azide in Et<sub>2</sub>O, dry it over MgSO<sub>4</sub>, filter, evaporate and recrystallise the residue from petroleum ether (b 20-40°) to give orange needles. Store it in a stoppered container at ~0°. The <sup>1</sup>H NMR has δ at 7.75 (m 1H) and 7.35 (m 2H) in CDCl<sub>3</sub>. [Hagedorn et al. *J Org Chem* **43** 2070 1978.]

***o*-Fluorophenol** [367-12-4] **M 112.1, m 16°, b 53°/14mm, 171-172°/714mm,  $d_4^{20}$  1.257,  $n_D^{20}$  1.514,  $pK^{25}$  8.70.** Pass *o*-fluorophenol at least twice through a gas chromatographic column for small quantities; otherwise fractionally distil it under reduced pressure. [Beilstein 6 I 97, 6 IV 770.]

***p*-Fluorophenoxyacetic acid** [405-79-8] **M 170.1, m 104.2-104.6°, 106°,  $pK^{25}$  3.13.** Crystallise the acid from EtOH or H<sub>2</sub>O. [Hayes & Branch *J Am Chem Soc* 65 1555 1943, Beilstein 6 III 971, 6 IV 776.]

**4-Fluorophenylacetic acid** [405-50-5] **M 154.1, m 86°, 94°, b 164°/2mm,  $pK^{25}$  4.22.** Crystallise it from heptane, but it is best purified by distillation at high vacuum. [Bergmann et al. *J Am Chem Soc* 78 6037 1956, Beilstein 9 III 2261, 9 IV 1672.]

**4-Fluorophenyl isocyanate** [1195-45-5] **M 137.1, b 55°/8mm,  $n_D^{20}$  1.514.** Purify the isocyanate by repeated fractionation through an efficient column. If its IR indicates that there is too much urea (in the presence of moisture the symmetrical urea is formed), then dissolve it in dry EtOH-free CHCl<sub>3</sub>, filter, evaporate and distil it. **It is a pungent LACHRYMATORY liquid, TOXIC.** [See Hardy *J Chem Soc* 2011 1934, and Hickinbottom *Reactions of Organic Compounds* Longmans p493 1957.]

**4-Fluorophenyl isothiocyanate** [1544-68-9] **M 153.2, m 24-26°, 26-27°, b 66°/2mm, 215°/atm, 228°/760mm,  $n_D^{20}$  1.6116.** A likely impurity is the symmetrical thiourea. Dissolve the isothiocyanate in dry CHCl<sub>3</sub>, filter and distil the residue in a vacuum. It can also be steam distilled, the oily layer is separated, dried over CaCl<sub>2</sub> and distilled *in vacuo*. *Bis-(4-fluorophenyl)thiourea* has **m 145°** (from aqueous EtOH). [Browne & Dyson *J Chem Soc* 3285 1931, Buu Hoi et al. *J Chem Soc* 1573 1955, Olander *Org Synth Coll Vol I* 448 1941].

***p*-Fluorophenyl-*o*-nitrobenzyl ether** [448-37-3] **M 247.2, m 62°.** Crystallise the ether from EtOH. [Jones *J Chem Soc* 1416 1938, Beilstein 6 III 1564, 6 IV 2608.]

***o*-Fluorotoluene** [95-52-3] **M 110.1, m -62°, b 114.4°/760mm,  $d_4^{20}$  1.005,  $n_D^{20}$  1.475.** Dry *o*-fluorotoluene with P<sub>2</sub>O<sub>5</sub> or CaSO<sub>4</sub> and fractionally distil it through a silvered vacuum-jacketed glass column with 1/8th-in glass helices. A high reflux ratio is necessary because of the closeness of the boiling points of the *o*-, *m*- and *p*-isomers [Potter & Saylor *J Am Chem Soc* 37 90 1915]. [Beilstein 5 H 290, 5 III 676, 5 IV 799.]

***m*-Fluorotoluene** [352-70-5] **M 110.1, m -87°, b 116.5°/760mm,  $d_4^{20}$  1.00,  $n_D^{27}$  1.46524.** Purify it as for *o*-fluorotoluene. [Beilstein 5 H 290, 5 III 676, 5 IV 799.]

***p*-Fluorotoluene** [352-32-9] **M 110.1, m -56°, 116.0°/760mm,  $d_4^{20}$  1.00,  $n_D^{20}$  1.46884.** Purify it as for *o*-fluorotoluene. [Beilstein 5 H 290, 5 III 677, 5 IV 799.]

**Formanilide** [103-70-8] **M 121.1, m 50°, b 166°/14mm, 216°/120mm,  $d_4^{20}$  1.14.** Crystallise formanilide from Et<sub>2</sub>O (**m 45.3°**), Et<sub>2</sub>O/petroleum ether (**m 46°**), petroleum ether (**m 47.6°**), ligroin/xylene, or distil it preferably under reduced pressure. [Beilstein 12 H 230, 12 II 135, 12 III 453.]

**Gallic acid (H<sub>2</sub>O) (3,4,5-trihydroxybenzoic acid)** [5995-86-8 (H<sub>2</sub>O), 149-91-7 (anhydrous)] **M 188.1, m 253°(dec),  $pK_1^{25}$  4.27,  $pK_2^{25}$  8.68.** Crystallise gallic from water. The *tri-O-acetyl* derivative has **m 172°** (from MeOH), and the *anilide* has **m 207°**(from EtOH). [Beilstein 10 H 470, 10 IV 1993.]

**Galvinoxyl [2,6-di-*tert*-butyl- $\alpha$ -(3,5-di-*tert*-butyl-4-oxo-2,5-cyclohexadiene-1-ylidene)-*p*-tolyl]oxy** [2370-18-5] **M 421.65, m 153.2-153.6°, 158-159°.** It is a stable free radical scavenger of short-lived free radicals with odd electrons on C or O. It is best prepared freshly by oxidation of 3,3',5,5'-tetra-*tert*-butyl-4,4'-dihydroxydiphenyl-methane [**m 154°**, 157.1-157.6°, obtained by gently heating for 10-15 minutes 2,6-di-*tert*-butylphenol, formaldehyde and NaOH in EtOH and recrystallised from EtOH (20g/100ml) as colourless plates, Karasch & Joshi *J Org Chem* 22 1435 1957, Bartlett et al. *J Am Chem Soc* 82 1756 1960 and 84 2596 1962]. Oxidation is carried out under N<sub>2</sub> with PbO<sub>2</sub> in Et<sub>2</sub>O or isooctane [Galvin A. Coppinger *J Am Chem Soc* 79 501 1957, Bartlett et al. above] or with alkaline potassium ferricyanide [Karasch & Joshi, above], whereby

Galvinoxyl separates as deep blue crystals and is recrystallised twice under  $N_2$  from  $*C_6H_6$  solution by suction evaporation at  $30^\circ$ . The VIS spectrum has  $\lambda_{max}$  at 407nm ( $\epsilon$  30,000), 431nm ( $\epsilon$  154,000) and weak absorption at 772.5nm, and IR has  $\nu_{max}$  at 1577 and  $2967cm^{-1}$ , and is estimated by iodometric titration. It is sensitive to  $O_2$  in the presence of  $OH^-$  ions and traces of strong acid in hydroxylic or hydrocarbon solvents. At  $62.5^\circ$  in a 0.62mM solution in  $*C_6H_6$  the radical decays with a first order  $k = 4 \times 10^{-8} \text{ sec}^{-1}$  (half life  $1.7 \times 10^{17} \text{ sec}$ ,  $\sim 200$  days) as observed by the change in OD at 550nm [see Green & Adam *J Org Chem* **28** 3550 1963].

**R(-)-Guaiaretic acid [guaiacic acid, 4,4'-(2,3-dimethyl-1-butene-(1,4-diyl)-bis-(2-methoxyphenol)]** [500-40-3] **M 328.4, m 99-100.5 $^\circ$ ,  $[\alpha]_D^{16} -91^\circ$  (c 1.1, EtOH),  $pK_{Est} \sim 10.0$** . Crystallise the acid once from petroleum ether (b  $60-80^\circ$ ), twice from 60% aqueous MeOH and finally from EtOH (m  $100-101^\circ$ ). Its UV has  $\lambda_{max}$  at 207 and 260nm ( $\log \epsilon$  4.64 and 4.28). [King et al. *J Chem Soc* 4011 1964, Schrecker & Hartwell *J Org Chem* **21** 381 1956.]

**Guaiacol (2-methoxyphenol)** [90-05-1] **M 124.1, m  $32^\circ$ , b  $106^\circ/24mm$ ,  $205^\circ/746mm$ ,  $pK^{25} 9.90$** . Crystallise guaiacol from  $*benzene$ /petroleum ether or distil it in a vacuum. [*Beilstein* **6** H 768, **6** IV 5563.]

**Guaiacol carbonate** [553-17-3] **M 274.3, m  $88.1^\circ$ ,  $89^\circ$** . Crystallise the carbonate from EtOH. It is estimated by bromination to the *monobromo guaiacol carbonate* m  $178^\circ$  (from EtOH) [Chernoff *J Am Chem Soc* **51** 3073 1929]. [*Beilstein* **6** H 776, **6** I 386, **6** II 784, **6** III 4233.]

**Guaiazuline (1,4-dimethyl-7-isopropylazulene)** [489-84-9] **M 198.3, m  $31.5^\circ$ , b  $153^\circ/7mm$ ,  $d^{25} 0.976$** . It forms blue-violet plates from EtOH. Also redistil it *in vacuo* until distillate solidifies. Its UV has  $\lambda_{max}$  at 284nm ( $\log \epsilon$  4.6, heptane). The *picrate* has m  $122^\circ$ (EtOH). [*Beilstein* **5** III 1677, **5** IV 1751.]

**Guanabenz (1-[2,6-dichlorobenzylideneamino]guanidine WY-8677)** [5051-62-7] **M 231.1, m  $227-229^\circ$ (dec),  $pK_{Est} \sim 9$** . Crystallise the guanidine from MeCN and store it in a  $CO_2$ -free atmosphere. The *monoHCl* [23256-5-0] has m  $178^\circ$ (dec) [solubility w% at  $25^\circ$  is 1.1( $H_2O$ ), 5(EtOH), 0.1(EtOAc), 0.06( $CHCl_3$ )]. It is an antihypertensive drug. [Holmes et al. *Drugs* **26** 212 1983.]

**Hexachlorobenzene** [118-74-1] **M 284.8, m  $230.2-231.0^\circ$ , b  $323-326^\circ/atm$** . Crystallise hexachlorobenzene repeatedly from  $*benzene$ . Dry it under vacuum over  $P_2O_5$ . [*Beilstein* **5** H 205, **5** IV 670.]

**Hexachlorophene (2,2'-methylenebis[3,4,6-trichlorophenol])** [70-30-4] **M 406.9, m  $166-167^\circ$ ,  $pK_1^{20} 5.41$ ,  $pK_2^{20} (30\% \text{ aqueous MeOH})$** . It forms needles from MeOH,  $C_2H_4Cl_2$ , or toluene. The *diacetate* has m  $175-176^\circ$  (EtOH). A disinfectant is also available in MeOH (5mg/l). [*Beilstein* **6** III 5407, **6** IV 6659.]

**Hexadecyl 3-hydroxynaphthalene-2-carboxylate** [531-84-0] **M 412.6, m  $73-74^\circ$ ,  $d^{135} 0.824$** . Recrystallise the ester from hot EtOH and sublime it in a vacuum. It is soluble in  $*C_6H_6$ , petroleum ether and AcOH. [Oshima & Hayashi *J Soc Chem Ind Jpn* **44** 821 1941, *Chem Abs* **42** 2108 1948, *Beilstein* **5** H 471, **5** I 227, **5** II 358, **5** III 1106, **5** IV 1208.]

**Hexaethylbenzene** [604-88-6] **M 246.3, m  $128.7-129.5^\circ$** . Crystallise this hydrocarbon from  $*C_6H_6$ , or  $*C_6H_6$ /EtOH. The *1,3,5-trinitrobenzene complex (1:1)* has m  $174^\circ$ (EtOH). [*Beilstein* **5** III 3038, **5** IV 1137.]

**Hexafluorobenzene** [392-56-3] **M 186.1, m  $5.1^\circ$ , b  $79-80^\circ/760mm$ ,  $d_4^{20} 1.61$ ,  $n_D^{20} 1.378$** . Main impurities are incompletely fluorinated benzenes. Purify it by standing in contact with oleum for 4 hours at room temperature, repeating until the oleum does not become coloured. Wash it several times with water, then dry it with  $P_2O_5$ . Finally purify it by repeated fractional crystallisation. [*Beilstein* **5** III 523, **5** IV 640.]

**Hexamethylbenzene** [87-85-4] **M 162.3, m  $165-165.5^\circ$ ,  $165.75^\circ$ , b  $264^\circ/atm$** . It is prepared, in an efficient fume cupboard, by passing a solution of phenol (100g, 1.06moles) in MeOH (1L) dropwise at a rate of 110ml/hour through a column (5.08cm diameter  $\times$  40.6cm long) packed with  $Al_2O_3$  (300g, 4-8 or 8-14 mesh) as

catalyst held at 370-380°, and arranged so that effluent is connected to a receiver that allows the exit gasses (CO, CH<sub>4</sub> and H<sub>2</sub>) into the fume cupboard. The yellow product (112-115g, 65-67%, m 134-145°) from the effluent which solidifies is washed with cold MeOH, and purified by recrystallisation from EtOH (50g/ 650ml) with 65-67% recovery, or from \*C<sub>6</sub>H<sub>6</sub> (50g/130ml) with 60% recovery to give pure colourless *hexamethylbenzene* **m 165-166°**. [Cullinane et al. *Org Synth Coll Vol IV* 520 1963.] It is also formed in >80% yields when *hexamethyl-Dewar benzene* [HMDB, 7641-77-2] comes into contact with anhydrous Al<sub>2</sub>Cl<sub>3</sub> in hot \*benzene (highly exothermic) [Schäfer & Hellmann *Angew Chem Internat Edn* 6 518 1967]. It has been purified by sublimation, then recrystallisation from absolute EtOH, \*benzene, EtOH/\*benzene or EtOH/cyclohexane. It has also been purified by zone melting. Dry it in a low vacuum over P<sub>2</sub>O<sub>5</sub>. It forms 1:1 addition compounds with 2,4,7-trinitrofluorenone (m 165-166.3°), 1,3,5-trinitrobenzene (m 173°), 2,4,6-trinitrotoluene (m 123°), 2-chloro-1,3,5-trinitrobenzene (m 151°), and 1:1 complexes with Ag<sup>+</sup>, Cr(CO)<sub>3</sub> and W(CO)<sub>3</sub>. Its FT-IR (Nujol) has  $\nu_{\max}$  at 1300.8, 1249.1, 1180.6, 1129.2, 1057.1, 994.9, 943.5, 794.7 and 772.0 cm<sup>-1</sup>; the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 2.22 (s, 18H, Me<sub>6</sub>); and the <sup>13</sup>C NMR (15MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 17.72 (6-Me C) and 132.89 (6-aromatic C). [*Beilstein* 5 H 450, 5 I 213, 5 II 341, 5 III 1038, 5 IV 1137.]

**Homophthalic acid** [89-51-0] **M 180.2, m 180-181°, 182-183°, 189-190°, (depends on heating rate) pK<sub>Est(1)</sub> ~3.5, pK<sub>Est(2)</sub> ~4.3.** Crystallise the acid from boiling H<sub>2</sub>O (25ml/g; 63ml/g at 20°), aqueous AcOH (**m 180°**) Dry it at 100°. The *S*-benzylisothiuronium salt has **m 155-156°** (from EtOH). [Price *Org Synth* 22 61 1942, Grummitt et al. *Org Synth Coll Vol III* 449 1955, *Beilstein* 9 H 857, 9 II 617, 9 III 4266, 9 IV 3343.] The *anhydride* [703-59-3] **M 162.1** has **m 141-142°**. [*Beilstein* 17/11 V 270.]

**Homoveratronic nitrile (3,4-dimethoxybenzyl nitrile)** [93-17-4] **M 177.2, m 62-64°, 68°, b 184°/20mm, 195-196°/2mm, 208°/atm.** Its solubility is 10% in MeOH, and it has been recrystallised from EtOH or MeOH. Purify it also by distillation followed by recrystallisation. [Price & Rogers *Org Synth Coll Vol III* 174 1955, Niederl & Ziering *J Am Chem Soc* 64 885 1952, Julian & Sturgis *J Am Chem Soc* 57 1126 1935, *Beilstein* 10 I 198.]

**Homoveratrylamine (2-[3,4-dimethoxyphenyl]ethylamine)** [120-20-7] **M 181.2, b 99.3-101.3°/0.5mm, 168-170°/15mm, d<sub>4</sub><sup>20</sup> 1.091, n<sub>D</sub><sup>20</sup> 1.5460, pK<sub>Est</sub> ~9.8.** Purify the amine by fractionation through an efficient column in an inert atmosphere as it is a relatively strong base. [Horner & Sturm *Justus Liebigs Ann Chem* 608 12819 1957, Jung et al. *J Am Chem Soc* 75 4664 1953.] The *hydrochloride* has **m 152°, 154°, 156°** (from EtOH, Me<sub>2</sub>CO or EtOH/Et<sub>2</sub>O), the *picrate* has **m 165-167°(dec)**, and the *4-nitrobenzoyl* derivative has **m 147°** [Buck *J Am Chem Soc* 55 2593 1933]. [*Beilstein* 13 H 800, 13 IV 2604.]

**Hordenine {4-[(2-dimethylamino)ethyl]phenol}** [539-15-1] **M 165.2, m 117-118°, b 173-174°/11mm, pK<sup>25</sup> 9.46 (OH).** Crystallise Hordenine from EtOH or H<sub>2</sub>O and sublime it at 105°/5mm. The *4,4'-dichlorodiphenyl disulfimide salt* has **m 145-146°** (from Me<sub>2</sub>CO/H<sub>2</sub>O) [Runge et al. *Chem Ber* 88 50 1955]. [*Beilstein* 13 H 626, 13 I 236, 13 II 356, 13 III 1640, 13 IV 1790.]

**4-Hydrazinobenzoic acid** [619-67-0] **M 152.2, m 217° (dec), pK<sup>25</sup> 4.13.** The acid crystallises from water. [*Beilstein* 15 H 631, 15 IV 1372.] The *hydrochloride* [24589-37-3] **M 188.6**, has **m 253°(dec)** [*Beilstein* 15 III 837].

**Hydrobenzamide [1-phenyl-*N,N'*-bis(phenylmethylene)-methanediamine]** [92-29-5] **M 298.4, m 101-102°, 107-108°.** Crystallise hydrobenzamide from absolute EtOH, petroleum ether (**m 107-108°**), \*C<sub>6</sub>H<sub>6</sub> (**m 103°**), or cyclohexane/\*benzene. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. [Pirrone *Gazz Chim Ital* 67 534 1937, *Beilstein* 7 H 215, 7 I 120, 7 II 166, 7 III 838.]

***dl*-Hydrobenzoin (1,2-diphenyl-1,2-ethanediol, *iso*-hydrobenzoin)** [492-70-6] **M 214.3, m 120°.** Crystallise the diol from Et<sub>2</sub>O/petroleum ether or H<sub>2</sub>O (**m 121-122°**) [*Beilstein* 6 H 1004, 6 I 490, 6 II 969, 6 III 5431.] The *R,R*-(+)- and *S,S*-(-)- enantiomers have **m 148.5-149.5°(dec)**, [ $\alpha$ ]<sub>D</sub><sup>20</sup> + and -94.0° (c 2, EtOH) [Eisenlohr & Hill *Chem Ber* 70 942 1937].

***meso*-Hydrobenzoin** [579-43-1] **M 214.3, m 139°, 139-140°.** Crystallise it from EtOH or water. [*Beilstein* 6

H 1003, 6 I 490, 6 II 967, 6 III 5429, 6 IV 6682.]

**Hydroquinone (1,4-dihydroxybenzene, quinol)** [123-31-9] M 110.1, m 175.4, 176.6°, pK<sub>1</sub><sup>20</sup> 9.91, p K<sub>2</sub><sup>20</sup> 11.56. Crystallise quinol from acetone, \*benzene, EtOH, EtOH/\*benzene, water or acetonitrile (25g in 30ml), preferably under nitrogen. Dry it under vacuum. [Wolfenden et al. *J Am Chem Soc* 109 463 1987, *Beilstein* 6 H 836, 6 IV 5712.]

***p*-Hydroxyacetophenone** [99-93-4] M 136.2, m 109°, 111°, 147-148°/3mm, pK<sub>1</sub><sup>25</sup> 8.01. Crystallise it from diethyl ether, aqueous EtOH or \*benzene/petroleum ether. [*Beilstein* 8 H 87, 8 IV 339.]

**3-Hydroxyanthranilic acid (2-amino-3-hydroxybenzoic acid)** [548-93-6] M 153.1, m >240°(dec), 245-265°(dec), 253-154°, 252-255°, λ<sub>max</sub> 298nm, log ε 3000 (0.1M HCl), pK<sub>1</sub><sup>20</sup> 2.7, pK<sub>2</sub><sup>20</sup> 5.19, pK<sub>3</sub><sup>20</sup> 10.12. Crystallise the acid from H<sub>2</sub>O or EtOH (m 254-255°). Sublime it at 170°/0.1mm. The *hydrochloride* has m 227° (from dilute HCl). [*Beilstein* 14 H 587, 14 III 1463, 14 IV 2071.] **Possible carcinogen.**

**5-Hydroxyanthranilic acid (2-amino-5-hydroxybenzoic acid)** [394-31-0] M 153.1, m 233-234°, 235-236°, 248°(dec), pK<sub>1</sub><sup>25</sup> 2.72, pK<sub>2</sub><sup>25</sup> 5.37, pK<sub>3</sub><sup>20</sup> 10.12. Crystallise the acid from water. The *benzamide* has m 240-242° (from AcOH). It is a **hypoglycemic agent**. [*Beilstein* 14 H 591, 14 II 357, 14 III 1468, 14 IV 2080.]

***m*-Hydroxybenzaldehyde** [100-83-4] M 122.1, m 108° pK<sub>1</sub><sup>25</sup> 8.98, pK<sub>2</sub><sup>25</sup> 15.81. Crystallise the aldehyde from water. [*Beilstein* 8 H 58, 8 IV 240.]

***p*-Hydroxybenzaldehyde** [123-08-0] M 122.1, m 115-116°, 117-119°, pK<sub>1</sub><sup>25</sup> 7.61. Crystallise it from water (containing some H<sub>2</sub>SO<sub>4</sub>). Dry it over P<sub>2</sub>O<sub>5</sub> under vacuum. [*Beilstein* 8 H 64, 8 IV 251.]

***m*-Hydroxybenzoic acid** [99-06-9] M 138.1, m 200.8°, pK<sub>1</sub><sup>25</sup> 4.08, pK<sub>2</sub><sup>25</sup> 9.98. Crystallise the hydroxyacid from absolute EtOH. [*Beilstein* 10 IV 315.]

***p*-Hydroxybenzoic acid** [99-96-7] M 138.1, m 213-214°, pK<sub>1</sub><sup>25</sup> 4.50, pK<sub>2</sub><sup>25</sup> 9.11. Crystallise the hydroxyacid from water. [*Beilstein* 10 IV 345.]

**2-Hydroxybenzyl alcohol (saligenine)** [90-01-7] M 124.1, m 86-87°, 87°, pK<sub>1</sub><sup>25</sup> 9.92. Crystallise saligenine from water or EtOH/\*C<sub>6</sub>H<sub>6</sub> (m 89°). [*Beilstein* 6 III 4537, 6 IV 5896.]

**3-Hydroxybenzyl alcohol** [620-24-6] M 124.1, m 71°, pK<sub>1</sub><sup>25</sup> 9.83. Crystallise the alcohol from \*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub>. [*Beilstein* 6 III 4545, 6 IV 5907.]

**4-Hydroxybenzyl alcohol** [623-05-2] M 124.1, m 122°, 127°, 132°, pK<sub>1</sub><sup>25</sup> 9.82. Crystallise the alcohol from H<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> (m 124°), \*C<sub>6</sub>H<sub>6</sub>/EtOH or ClCH<sub>2</sub>CH<sub>2</sub>Cl (m 122°). [*Beilstein* 6 III 4546, 6 IV 5909.]

**2-Hydroxybiphenyl** [90-43-7] M 170.2, m 56°, b 145°/14mm, 275°/760mm, pK<sub>1</sub><sup>20</sup> 10.01. Crystallise it from petroleum ether. [*Beilstein* 6 IV 4579.]

**4-Hydroxybiphenyl (4-phenylphenol)** [92-69-3] M 170.2, m 164-165°, b 305-308°/760mm, pK<sub>1</sub><sup>23</sup> 9.55. Crystallise the phenol from aqueous EtOH, \*C<sub>6</sub>H<sub>6</sub>, and dry it in a vacuum over CaCl<sub>2</sub> [Buchanan et al. *J Am Chem Soc* 108 7703 1986]. [*Beilstein* 6 IV 4600.]

***trans*-4-Hydroxycinnamic acid (*p*-coumaric acid)** [501-98-4] M 164.2, m 210-213°, 214-215°, 215°, pK<sub>1</sub><sup>25</sup> 4.64, pK<sub>2</sub><sup>25</sup> 9.45. Crystallise *p*-coumaric acid from H<sub>2</sub>O (charcoal). It forms needles from concentrated aqueous solutions as the *anhydrous acid*, but from hot dilute solutions the *monohydrate acid* separates on slow cooling. The acid (33g) has been crystallised from 2.5L of H<sub>2</sub>O (1.5g charcoal) yielding 28.4g of recrystallised acid, m 207°. It is insoluble in \*C<sub>6</sub>H<sub>6</sub> or petroleum ether. The UV in 95% EtOH has λ<sub>max</sub> at 223 and 286nm (ε 14,450 and 19000 M<sup>-1</sup>cm<sup>-1</sup>). [UV Wheeler & Covarrubias *J Org Chem* 28 2015 1963, Corti *Helv Chim Acta* 32 681 1949, *Beilstein* 10 IV 1005.]

**3-(4-Hydroxy-3,5-dimethoxyphenyl)acrylic acid (sinapinic acid)** [530-59-6] **M 234.1, m 204-205°(dec), pK<sub>Est(1)</sub> ~4.6, pK<sub>Est(2)</sub> ~9.3.** Crystallise it from water. [Beilstein 10 H 508, 10 IV 2104.]

**4-Hydroxydiphenylamine** [122-37-2] **M 185.2, m 72-73°, pK<sub>Est</sub> ~10.0.** Crystallise the amine from chlorobenzene/petroleum ether, pentane (**m 72°**) or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (**m 70°**). [Beilstein 13 III 1019, 13 IV 1052.]

**2-Hydroxy-4-(n-dodecyloxy)benzophenone** [2985-59-3] **M 382.5, m 50-52°, pK<sub>Est</sub> ~7.1.** Recrystallise it from *n*-hexane and then 10% (v/v) EtOH in acetonitrile [Valenty et al. *J Am Chem Soc* 106 6155 1984].

**4-Hydroxyindane** [1641-41-1] **M 134.2, m 49-50°, b 120°/12mm, pK<sup>25</sup> 10.32.** Crystallise 4-hydroxyindane from petroleum ether, pentane (**m 50-50.5°**) or \*C<sub>6</sub>H<sub>6</sub> (**m 39.5-40°**). It has UV with λ<sub>max</sub> at 277nm (cyclohexane). The *acetyl* derivative has **m 30-32°** (from EtOH), **b 127°/14mm** and the *3,5-dinitrobenzoyl* derivative has **m 114°**. [Dallacker et al. *Chem Ber* 105 2568 1972, Beilstein 6 III 2427, 6 IV 3827.]

**5-Hydroxyindane** [1470-94-6] **M 134.2, m 55°, b 127°/14mm, 255°/760mm, pK<sup>23</sup> 10.24.** Crystallise 5-hydroxyindane from petroleum ether (**m 56°**) or pentane (**m 59-60°**). It has UV with λ<sub>max</sub> at 283.5nm (cyclohexane). The *3,5-dinitrobenzoyl* derivative has **m 156°**. [Beilstein 6 III 2428, 6 IV 3829.]

**4-Hydroxy-3-methoxyacetophenone (apocynin)** [498-02-2] **M 166.2, m 115°, pK<sub>Est</sub> ~7.9.** Crystallise apocynin from water, or EtOH/petroleum ether. [Beilstein 8 IV 1814.]

**trans-4-Hydroxy-3-methoxycinnamic acid (ferulic acid)** [537-98-4; 1135-24-6] **M 194.2, m 174°, pK<sub>1</sub><sup>25</sup> 4.58, pK<sub>2</sub><sup>25</sup> 9.39.** Crystallise ferulic acid from H<sub>2</sub>O. [Beilstein 10 H 436, 10 IV 1776.]

**4-Hydroxy-2-methylazobenzene (3-methyl-4-phenylazophenol)** [1435-88-7] **M 212.2, m 100-101°, 112°, pK<sub>Est</sub> ~9.5.** Crystallise the phenol from hexane. [Beilstein 16 II 61, 16 IV 195.]

**4-Hydroxy-3-methylazobenzene (2-methyl-4-phenylazophenol)** [621-66-9] **M 212.2, m 125-126°.** Crystallise the phenol from hexane or petroleum ether (**m 130°**). [Beilstein 16 II 59, 16 III 104, 16 IV 193.]

**3-Hydroxy-4-methylbenzaldehyde** [57295-30-4] **M 136.1, m 73°, pK<sub>Est</sub> ~10.2.** Crystallise it from water or \*C<sub>6</sub>H<sub>6</sub> (**m 71-71°**). The *O-methyl ether* has **m 45-46°** (from Et<sub>2</sub>O/hexane). [Sedgwick & Allott *J Chem Soc* 2820 1923, Flitsch et al. *Justus Liebigs Ann Chem* 1413 1985, Beilstein 8 H 100, 8 II 103, 8 IV 368.]

**5-Hydroxy-2-methyl-1,4-naphthaquinone (plumbagin)** [481-42-5] **M 188.2, m 78-79°, pK<sub>Est(1)</sub> ~9.5, pK<sub>Est(2)</sub> ~11.0.** It crystallises in yellow needles from aqueous EtOH. It is soluble in organic solvents, it is steam volatile and it sublimes on heating in a vacuum. [Fieser & Dunn *J Am Chem Soc* 58 572 1936, Beilstein 8 III 2576, 8 IV 2376.]

**6-Hydroxy-2-methyl-1,4-naphthaquinone** [633-71-6] **M 188.2, pK<sub>Est</sub> ~10.0.** Crystallise the naphthaquinone from aqueous EtOH and sublime it in a vacuum.

**2-Hydroxy-1-naphthaldehyde** [708-06-5] **M 172.2, m 82°, b 139-142°/4mm, 192°/27mm, pK<sup>25</sup> 8.27 (50% aqueous EtOH).** Crystallise the aldehyde from EtOH (1.5ml/g), aqueous EtOH, aqueous AcOH (**m 84°**), EtOAc or H<sub>2</sub>O. [Russell & Lockhart *Org Synth Coll Vol III* 463 1955, Beilstein 8 H 143, 8 I 564, 8 II 171, 8 III 1108, 8 IV 1160.]

**2-Hydroxy-1-naphthaleneacetic acid** [10441-45-9] **M 202.2, m 147-148°, pK<sub>Est(1)</sub> ~4.2, pK<sub>Est(2)</sub> ~8.3.** Crystallise the acid from EtOH/water (1:9, v/v, activated charcoal), H<sub>2</sub>O (**m 157°**) or xylene (**m 147°**). Dry it under vacuum, over silica gel, in the dark. Store it in the dark at -20° [Gafni et al. *J Phys Chem* 80 898 1976]. It forms a cyclic lactone (**m 107°**) readily. [Beilstein 10 II 218, 10 III 1102, 10 IV 1201.]

**6-Hydroxy-2-naphthalenepropionic acid** [553-39-9] M 216.2, m 180-181°, pK<sub>Est(1)</sub> ~4.6, pK<sub>Est(2)</sub> ~9.0. Crystallise the acid from aqueous EtOH or aqueous MeOH. [Ormaney & Horeau *Bull Soc Chim Fr* 962 1955, *Beilstein* 10 III 1113.]

**3-Hydroxy-2-naphthalide (Naphthol AS)** [92-77-3] M 263.3, m 248.0-248.5°, CI 37505. Crystallise it from xylene or AcOH which forms plates m 243-244° [Schnopper et al. *Anal Chem* 31 1542 1959]. [*Beilstein* 12 H 505.]

**3-Hydroxy-2-naphtho-4'-chloro-*o*-toluidide** [92-76-2] M 311.8, m 243.5-244.5°. Crystallise it from xylene [Schnopper et al. *Anal Chem* 31 1542 1959].

**3-Hydroxy-2-naphthoic-1'-naphthylamide** [123-68-3] M 314.3, m 217-.5-218.0°. Crystallise the amide from xylene [Schnopper et al. *Anal Chem* 31 1542 1959].

**3-Hydroxy-2-naphthoic-2'-naphthylamide** [136-64-8] M 305.3, m 243.5-244.5°, and other naphthol AS derivatives. Crystallise it from xylene [Schnopper et al. *Anal Chem* 31 1542 1959].

**2-Hydroxy-1,4-naphthoquinone (Lawsone B, Neutral Orange 6, tautomeric with 4-hydroxy-1,2-naphthoquinone)** [83-72-7] M 174.2, m 192°(dec), CI 75480 pK<sub>1</sub><sup>25</sup> -5.6 (C=O protonation), pK<sub>2</sub><sup>25</sup> 2.38, pK<sub>3</sub><sup>25</sup> 4.00 (phenolic OH). Crystallise Lawsone B from \*C<sub>6</sub>H<sub>6</sub> or AcOH (m 192.5°, 195-196°). It sublimes in a vacuum (m 194°). It has UV with λ<sub>max</sub> at 455nm (aqueous NaOH). [*Beilstein* 8 H 300, 8 I 635, 8 II 344, 8 III 2543, 8 IV 2360.]

**5-Hydroxy-1,4-naphthoquinone (Juglone)** [481-39-0] M 174.2, m 155°, 164-165°, pK<sup>25</sup> 8.7. Crystallise Juglone from \*benzene/petroleum ether or petroleum ether. [*Beilstein* 8 III 2558, 8 IV 2368.]

**6-Hydroxy-2-naphthyl disulfide** [6088-51-3] M 350.5, m 221-222°, 226-227°, pK<sub>Est</sub> ~9.0. It crystallises as leaflets from AcOH and is slightly soluble in EtOH, and AcOH, but is soluble in \*C<sub>6</sub>H<sub>6</sub> and in alkalis to give a yellow solution. [Zincke & Dereser *Chem Ber* 51 352 1918.] The *acetoxy* derivative has m 198-200° (from AcOH or dioxane/MeOH), and the *diacetyl* derivative has m 167-168° (from AcOH). A small amount of impure disulfide can be purified by dissolving it in a small volume of Me<sub>2</sub>CO and adding a large volume of toluene, filter rapidly and concentrate to one-third of its volume. The hot toluene solution is filtered rapidly from any tarry residue, and crystals separate on cooling. Recrystallisation from hot acetic acid gives crystals with m 220-223° [Barrett & Seligman *Science* 116 323 1952]. [*Beilstein* 6 I 481.]

**2-Hydroxy-5-nitrobenzyl bromide (Koshland's reagent)** [772-33-8] M 232.0, m 147°, pK<sub>Est</sub> ~8.0. Crystallise the bromide from \*benzene or \*benzene/ligroin. It is slightly soluble in EtOH, soluble in \*C<sub>6</sub>H<sub>6</sub> and AcOH, and very soluble in ligroin. [*Beilstein* 6 H 367.]

**2-Hydroxyphenylacetic acid** [614-75-5] M 152.2, m 148-149°, 152-153°, b 240-243°/760mm, pK<sub>Est(1)</sub> ~4.3, pK<sub>Est(2)</sub> ~10.1. Crystallise the acid from ether or chloroform (m 147°, m from latter solvent is always lower). [*Beilstein* 10 H 187, 10 I 81, 10 II 112, 10 III 422, 10 IV 536.]

**3-Hydroxyphenylacetic acid** [621-37-4] M 152.2, m 137°, pK<sub>Ext(1)</sub> ~4.3, pK<sub>Ext(2)</sub> ~10. Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub>/ligroin or EtOAc/cyclohexane (m 131-132°). [*Beilstein* 10 H 189, 10 I 82, 10 II 112, 10 III 428, 10 IV 541.]

**4-Hydroxyphenylacetic acid** [156-38-7] M 152.2, m 150-151°, 152°, pK<sub>1</sub> 4.28, pK<sub>2</sub> 10.1. Crystallise the acid from water or Et<sub>2</sub>O/petroleum ether. The *p*-bromophenacyl ester has m 117° (from EtOH). [*Beilstein* 10 II 112, 10 III 430, 10 IV 543.]

***N*-(4-Hydroxyphenyl)-3-phenylsalicylamide** [550-57-2] M 305.3, m 183-184°, pK<sub>Est</sub> ~9.5. Crystallise the amide from aqueous MeOH. [*Beilstein* 13 IV 224.]

**L-2-Hydroxy-3-phenylpropionic acid (3-phenyl lactic acid)** [20312-36-1] **M 166.2, m 125-126°**,  $[\alpha]_{\text{D}}^{12}$  -18.7° (EtOH),  $[\alpha]_{\text{D}}^{20}$  -22° (c 1, H<sub>2</sub>O), **pK** see below. Crystallise the acid from water, MeOH, EtOH or \*benzene. [Beilstein 10 IV 653.]

**dl-2-Hydroxy-3-phenylpropionic acid** [828-01-3] **M 166.2, m 97-98°**, **b 148-150°/15mm**, **pK<sub>Est</sub> ~3.7**. Crystallise the propionic acid from \*benzene or chloroform. [Beilstein 10 IV 653.]

**3-p-Hydroxyphenylpropionic acid (phloretic acid)** [501-97-3] **M 166.2, m 129-130°**, **131-133°**, **pK<sub>Est(1)</sub> ~4.7**, **pK<sub>Est(2)</sub> ~10.1**. Crystallise phloretic acid from ether or H<sub>2</sub>O. [Beilstein 10 IV 631.]

**p-Hydroxyphenylpyruvic acid** [156-39-8] **M 180.2, m 220°(dec)**, **pK<sub>Est</sub> ~2.3**. Crystallise it three times from 0.1M HCl/EtOH (4:1, v/v) immediately before use [Rose & Powell *Biochem J* 87 541 1963], or from Et<sub>2</sub>O. The *3,4-dinitrophenylhydrazone* has **m 178°**. [Beilstein 10 IV 3630.]

**N-Hydroxyphthalimide** [524-38-9] **M 163.1, m 230°**, **~235°(dec)**, **237-240°**, **pK<sup>30</sup> 7.0**. Dissolve the imide in H<sub>2</sub>O by adding Et<sub>3</sub>N to form the salt and while hot, acidify, cool and pour into a large volume of H<sub>2</sub>O. Filter off the solid, wash it with H<sub>2</sub>O and dry it over P<sub>2</sub>O<sub>5</sub> in a vacuum. [Nefken & Teser *J Am Chem Soc* 83 1263 1961, Fieser 1 485 1976, Nefkens et al. *Rec Trav Chim Pays Bas* 81 683 1962] The *O-acetyl* derivative has **m 178-180°** (from EtOH). [Beilstein 21/11 V 100.]

**4'-Hydroxypropiophenone** [70-70-2] **M 150.2, m 149°**, **b 140-145°/0.5mm**, **pK<sup>25</sup> 8.05**. Crystallise the phenone from H<sub>2</sub>O (**m 149.8-150.2°**) or EtOH (**m 147°**). The *benzoyl derivative* has **m 117°**, and the *semicarbazone* has **m 183°** (EtOH). [Beilstein 8 H 102, 8 II 104, 8 III 379.]

**trans-4-Hydroxystilbene** [6554-98-9] **M 196.3, m 189°**. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>, MeOH (**m 186-187°**) or acetic acid. [Beilstein 6 H 693, 6 II 657, 6 III 3497, 6 4855.]

**Hydroxy(tosyloxy)iodobenzene [phenyl(hydroxy)tosyloxyiodine, hydroxy(4-methylbenzenesulfonato-O)phenyliodine, Koser's reagent]** [27126-76-7] **M 392.2, m 134-136°**, **135-138°**, **134-136°**, **136-138.5°**. Possible impurities are tosic acid (removed by washing with Me<sub>2</sub>CO) and acetic acid (removed by washing with Et<sub>2</sub>O). It is purified by dissolving in the minimum volume of MeOH, adding Et<sub>2</sub>O to cloud point and setting aside for the prisms to separate [Koser & Wettach *J Org Chem* 42 1476 1977, NMR: Koser et al. *J Org Chem* 41 3609 1976]. It has also been crystallised from CH<sub>2</sub>Cl<sub>2</sub> (needles, **m 140-142°**) [Neiland & Karele *J Org Chem, USSR (Engl Transl)* 6 889 1970].

**9-Hydroxytriptycene** [73597-16-7] **M 270.3, m 245-246.5°**. Crystallise it from \*benzene/petroleum ether. Dried it at 100° in a vacuum [Bartlett & Greene *J Am Chem Soc* 76 1088 1954, Imashiro et al. *J Am Chem Soc* 109 729 1987].

**Hypericin (hypericum, 4,5,7,4',5',7'-hexahydroxy-2',2'-dimethylnaphthodianthrone)** [548-04-9] **M 504.4, m 320°(dec)**. Crystallise hypericin from pyridine by addition of methanolic HCl. [Beilstein 8 IV 3761.]

**Ibuprofen [(S+) and (R-) 4-isobutyl- $\alpha$ -methylphenylacetic acid, Brufen, Nurofen, Motrui]** [(S+) 51146-56-6, (R-) 51146-57-7] **M 206.3, m 52-53°**,  $[\alpha]_{\text{D}}^{20}$  +59° (c 2, EtOH). Crystallise the (+) and (-) acids from EtOH or aqueous EtOH. The *racemate*, which crystallises from petroleum ether with **m 75-77°**, is sparingly soluble in H<sub>2</sub>O and has IR (film) with  $\nu_{\text{max}}$  at 1705 (C=O), 2300—3700 (OH broad) cm<sup>-1</sup>. It is used as a non-steroidal anti-inflammatory. [Shiori et al. *J Org Chem* 43 2936 1978, Kaiser et al. *J Pharm Sci* 65 269 1976, *J Pharm Sci* 81 221 1992, Freer *Acta Cryst (C)* 49 1378 1993 for the (S+)-enantiomer.]

**1,3-Indandione** [606-23-5] **M 146.2, m 129-132°**, **pK<sup>18</sup> 7.2 (1% aqueous EtOH)**. Recrystallise it from EtOH or \*C<sub>6</sub>H<sub>6</sub>. In dilute alkali it gives a deep yellow solution of the enol. [Bernasconi & Paschalis *J Am Chem Soc* 108 2969 1986]. [Beilstein 7 IV 2344.]

**Indane (indan, hydrindene, 1,2-trimethylenebenzene)** [496-11-7] **M 118.1, f -51.4°, b 79°/29mm, 177-179.5°/760mm,  $d_4^{20}$  0.960,  $n_D^{20}$  1.536.** Shake indane with conc  $H_2SO_4$ , then water, dry ( $Na_2SO_4$ ) and fractionally distil it. [Beilstein 5 H 486, 5 I 234, 5 II 376, 5 III 1200, 5 IV 1371.]

**Indene** [95-13-6] **M 116.2, f -1.5°, m -5° to -3°, b 114.5°/100mm, 181-182°/atm,  $d_4^{20}$  0.994,  $n_D^{20}$  1.5763.** Shake indene with 6M HCl for 24 hours (to remove basic nitrogenous material), then reflux it with 40% NaOH for 2 hours (to remove benzonitrile). Fractionally distil, then fractionally crystallise it by partial freezing. The higher-melting portion is converted to its sodium salt by adding a quarter of its weight of sodamide under nitrogen and stirring for 3 hours at 120°. Unreacted organic material is distilled off at 120°/1mm. The sodium salts are hydrolysed with water, and the organic fraction is separated by steam distillation, followed by fractional distillation. Before use, the distillate is passed, under nitrogen, through a column of activated silica gel. It turns yellow in air as it readily oxidises and polymerises. Store it in the presence of *tert*-butylcatechol (50-100ppm) as antioxidant. [Russell *J Am Chem Soc* 78 1041 1956, Beilstein 5 IV 1532.]

**2-Iodoaniline** [615-43-0] **M 219.0, m 60-61°,  $pK^{25}$  2.54.** Distil 2-iodoaniline with steam and crystallise it from \*benzene/petroleum ether. The *N*-acetyl derivative has **m 110°.** [Beilstein 12 IV 1542.]

**4-Iodoaniline** [540-37-4] **M 219.0, m 62-63°,  $pK^{25}$  3.81.** Crystallise it from petroleum ether (b 60-80°) by refluxing, then cool it in an ice-salt bath freezing mixture. Dry it in air. *Alternatively*, crystallise it from EtOH and dry it *in vacuo* for 6 hours at 40° [Edidin et al. *J Am Chem Soc* 109 3945 1987]. The *N*-acetyl derivative has **m 184°** (from MeOH). [Beilstein 12 IV 1544.]

**4-Iodoanisole** [696-62-8] **M 234.0, m 51-52°, b 133-133.3°/25mm, 139°/35mm, 237°/726mm.** Crystallise 4-iodoanisole from aqueous EtOH and/or distil it under vacuum. [Beilstein 6 H 208, 6 I 109, 6 II 199, 6 III 744, 6 IV 1075.]

**Iodobenzene** [591-50-4] **M 204.0, b 63-65°/10mm, 188°/atm,  $d_4^{20}$  1.829,  $n_D^{25}$  1.6169.** Wash it with dilute aqueous  $Na_2S_2O_3$ , then water. Dry it with  $CaCl_2$  or  $CaSO_4$ , decolourise with charcoal and distil it under reduced pressure then store it with mercury or silver powder to stabilise it. [Beilstein 5 IV 688.]

***o*-Iodobenzoic acid** [88-67-5] **M 248.4, m 161-162°, 161.6-162°, 162°, 162.7-163.5°, 164°,  $pK^{20}$  2.93,  $pK^{25}$  2.84 ( $H_2O$ ),  $pK^{20}$  2.93 (1% EtOH),  $pK$  -7.78 ( $H_2SO_4$ ).** Crystallise the acid repeatedly from water (charcoal), EtOH, aqueous EtOH, aqueous  $Me_2CO$ ,  $Me_2CO/Et_2O$  or  $*C_6H_6$ . Sublime it under vacuum at 100°. *o*-Iodobenzoic acid is prepared by the following procedure: anthranilic acid (31.9g, 233mmol, see [118-92-3]) is dissolved in dilute  $H_2SO_4$  (50g, 27ml, 98% acid, 500mmol, in 250ml of  $H_2O$ ), cooled to 0°; and while stirring  $NaNO_2$  (20g, 290mmol) in  $H_2O$  (40ml) is added dropwise keeping the temperature below 10° (add crushed ice to the mixture if necessary). Diazotisation is complete when starch-iodide paper wetted with a drop of the mixture turns blue or brown due to the presence of excess of  $HNO_2$ . A solution of KI (60g, 360mmol) in  $H_2O$  (100ml) is gradually added with stirring, the mixture is set aside for 1 hour then heated at ~100° until effervescence ( $N_2$  liberated) ceases. The mixture is cooled in ice, the brown iodo-acid (~92%) is filtered off, washed well with  $H_2O$  and dried *in vacuo*. The brown colour is not easily removed by recrystallisation so the acid is converted into the ethyl ester e.g. by Fischer and Speier's method [Fischer & Speier *Chem Ber* 28 2250 1895]. Dry HCl gas is bubbled through EtOH (~200ml) cooled in ice-water until the increase in weight is ~4g. The crude iodobenzoic acid is added to it and the mixture is boiled under reflux for 2 hours, or until an aliquot poured into  $H_2O$  does not deposit a solid. The mixture is poured into cold  $H_2O$ , the heavy oil that separated is extracted into  $Et_2O$ , the extract is washed with  $H_2O$ , saturated aqueous  $NaHCO_3$  (care-effervescence), brine, dried ( $CaCl_2$ ), filtered, evaporated and the residual oil is distilled to give a 73% yield of pure *ethyl o-iodobenzoate* **M 276.0** [1829-28-3], **b 151-152°/12-13mm (163-165°/23mm, 275°/atm).** This ester is dissolved in 0.5 *N* alcoholic KOH (from 7g of KOH dissolved in 7ml of  $H_2O$  and diluted to 250ml with EtOH) in the ratio of 1g:25ml, and refluxed for 30 minutes (or until a drop added to  $H_2O$  dissolves completely), poured it into  $H_2O$ , and the clear solution of the potassium salt is acidified to pH <0 when pure iodobenzoic acid (theoretical yield) separates. It is collected, washed with cold  $H_2O$ , and dried *in vacuo*. [Baker et al. *J Chem Soc* 3721 1965, Cohen & Raper *J Chem Soc* 1271 1904.] The FT-IR (Nujol) has  $\nu_{max}$  at 3061.7, 2646.1, 1683.9, 1582.4, 1405.0,

1269.8, 1015.8, 739.0 and 679.7  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR ( $\text{DMSO}_d_6$ , TMS) has  $\delta$  at 7.25 (t of d, 1H,  $J = 7$  Hz and 1.5 Hz, H-5), 4.85 (t, 1H,  $J = 7$  Hz, H-4), 7.72 (d of d, 1H,  $J = 7$  and 1.5 Hz, H-6) and 8.01 (d, 1H,  $J = 7$  Hz, H-3); and the  $^{13}\text{C}$  NMR ( $\text{DMSO}_d_6$ , TMS) has  $\delta$  at 167.93, 140.32, 136.72, 132.25, 129.88 127.97 and 93.92.

The *acid chloride* has **M 266.4** [609-67-6], **m 35-40°**, **b 135°/19mm (159°/27mm)**, and the *amide* has **M 247.0** [3930-83-4], **m 183°** (needles from  $\text{H}_2\text{O}$ ). The *4-bromobenzylisothiuronium salt* has **m 154°** (from EtOH), and the *4-chlorobenzylisothiuronium salt* has **m 162°** (from dioxane). [*Beilstein* 9 H 363, 9 I 148, 9 II 239, 9 III 1432, 9 IV 1030.]

***m*-Iodobenzoic acid** [618-51-9] **M 248.4**, **m 185-186°**, **186.6-186.8°**, **189°**, **189-189.2°**, **pK<sup>25</sup> 3.85**, **pK<sup>50</sup> 3.97**, **pK -7.64 ( $\text{H}_2\text{SO}_4$ )**. Crystallise the acid repeatedly from water, EtOH, aqueous EtOH,  $\text{Me}_2\text{CO}$  or aqueous AcOH. Sublime it under vacuum at 100°. *Ethyl m-iodobenzoate* **M 276.0** [58313-23-8] has **b 150.5°/15mm (165-166°/24mm)**, the *m-acid chloride* has **M 266.4** [1711-10-0], **b 104-105°/1mm (159-150°/23mm)**, and the *m-amide* has **M 247.0** [10388-19-9], **m 186°** (needles from  $\text{H}_2\text{O}$ ). The *4-bromobenzylisothiuronium salt* has **m 152°** (from EtOH), and the *4-chlorobenzylisothiuronium salt* has **m 154°** (from dioxane). [*Fieser Org Synth Coll Vol II* 353 1948, *Beilstein* 9 H 365, 9 I 148, 9 II 240, 9 III 1437, 9 IV 1033.]

***p*-Iodobenzoic acid** [619-58-9] **M 248.4**, **m 269-270°**, **271-272°**, **272.5°**, **273°**, **pK<sup>25</sup> 4.00**, **pK -7.50 ( $\text{H}_2\text{SO}_4$ )**. Crystallise the acid repeatedly from water, aqueous EtOH, aqueous  $\text{Me}_2\text{CO}$ , aqueous AcOH. Sublime it under vacuum at 200°/0.4mm. *Ethyl p-iodobenzoate* **M 276.0** [51934-41-9] has **b 154°/15mm**, the *p-acid chloride* has **M 266.4** [1711-02-0] has **m 83° (77-78°)** (needles from  $\text{Et}_2\text{O}$ ), **b 126°/9mm (163-164°/32mm)**, and the *p-amide* has **M 247.0** [3956-07-8], **m 217.6° (209°)** (needles from  $\text{H}_2\text{O}$ ). The *4-bromobenzylisothiuronium salt* has **m 181°** (from EtOH), and the *4-chlorobenzylisothiuronium salt* has **m 177°** (from dioxane). [*Whitmore & Woodward Org Synth Coll Vol II* 325 1948, *Beilstein* 9 H 366, 9 I 149, 9 II 240, 9 III 1442, 9 IV 1035.]

**4-Iodobiphenyl** [1591-31-7] **M 280.1**, **m 113.7-114.3°**, **b 207°/28mm**. Crystallise 4-iodobiphenyl from EtOH/ $\text{C}_6\text{H}_6$ , and dry it *in vacuo* over  $\text{P}_2\text{O}_5$ . [*Beilstein* 5 H 581, 5 I 273, 5 II 486, 5 III 1748, 5 IV 1821.]

**1-Iodo-2,4-dinitrobenzene** [709-49-9] **M 294.0**, **m 88°**. Crystallise it from EtOAc. [*Beilstein* 5 H 270.]

**1-Iodo-4-nitrobenzene** [636-98-6] **M 249.0**, **m 171-172°**. Precipitate it from acetone by addition of water, followed by recrystallisation from EtOH. [*Beilstein* 8 H 523, 8 II 191, 8 III 623, 8 IV 743.]

***o*-Iodophenol** [533-58-4] **M 280.1**, **m 43°**, **pK<sup>25</sup> 8.51**. Crystallise 2-iodophenol from  $\text{CHCl}_3$  or diethyl ether. The *acetate* has **m 65-66°** (from MeOH). [*Beilstein* 6 H 208, 6 II 198, 6 III 774, 6 IV 1074.]

***p*-Iodophenol** [540-38-5] **M 280.1**, **m 94°**, **138-140°/5mm**, **pK<sup>25</sup> 9.30**. Crystallise 4-iodophenol from petroleum ether (b 80-100°) or distil it *in vacuo*. If the material has a brown or violet colour, then dissolve it in  $\text{CHCl}_3$ , shake it with 5% sodium thiosulfate solution until is colourless. Dry ( $\text{Na}_2\text{SO}_4$ ) the organic layer, filter, evaporate and distil the residue *in vacuo*. [*Dains & Eberly Org Synth Coll Vol II* 355 1943, *Beilstein* 6 IV 1074.]

**5-Iodosalicylic acid (2-hydroxy-5-iodobenzoic acid)** [119-30-2] **M 264.0**, **m 197°** **pK<sub>1</sub><sup>25</sup> 2.65**, **pK<sub>2</sub><sup>25</sup> 13.05**. Crystallise the acid from water. [*Beilstein* 10 H 112.]

***o*-Iodosobenzoic acid** [304-91-6] **M 264.0**, **m >200°**, **pK<sub>Est</sub><sup>25</sup> ~2.6**. Crystallise the acid from EtOH and dry it *in vacuo*. [*Beilstein* 9 H 363.]

***p*-Iodotoluene** [624-31-7] **M 218.0**, **m 35°**, **b 211-212°**. Crystallise 4-iodotoluene from EtOH and/or distil it. [*Beilstein* 5 IV 840.]

**Isonitrosoacetophenone (phenylglyoxaldoxime)** [532-54-7] **M 149.2**, **m 126-128°**. Crystallise it from water. [*Beilstein* 7 IV 2132.]

**Isophthalic acid (benzene-1,3-dicarboxylic acid)** [121-91-5] **M 166.1**, **m 345-348°**, **pK<sub>1</sub><sup>25</sup> 3.70**, **pK<sub>2</sub><sup>25</sup> 4.60**.

Crystallise the acid from aqueous EtOH. [Beilstein 9 IV 3292.]

**4,4'-Isopropylidenediphenol (Bisphenol A, 2,2-bis[4-hydroxyphenol]propane)** [80-05-7] **M 228.3, m 158°, pK<sub>Est</sub> ~10.3.** Crystallise bisphenol from acetic acid/water (1:1). It is used for making polycarbonate bottles and leaches out slowly on heating. It is a known "estrogenic chemical" shown to disrupt chemical signalling in the complex network of glands, hormones and cell receptors which make up the endocrine system. It causes low sperm count and damages the ecosystem by the feminisation of fish, reptiles and birds. [cf Chapter 1, p 3, Beilstein 6 IV 6717.]

**N-Isopropylidene-N'-2-nitrobenzenesulfonyl hydrazine (IPNSBH, isopropylidenehydrazide o-nitrobenzenesulfonic acid)** [6655-27-2] **M 257.3, m 132-135°, 139-140°.** IPNSBH is more stable than NBSH (*vide infra*) from which it is readily prepared by dissolving it (1.14g, 1 equivalent) in Me<sub>2</sub>CO (8ml) at 0°, and after 30 minutes the solvent is removed *in vacuo*, the residue is re-dissolved in Me<sub>2</sub>CO (4ml), and added slowly into 150ml of hexanes. The fine powder is filtered off, rinsed with hexanes (2 × 5ml) and dried *in vacuo* to give IPNSBH as a white solid (1.20g, 89%; TLC has R<sub>F</sub> 0.7, 66% EtOAc in hexanes). Its FTIR (thin film) has ν<sub>max</sub> at 1177s, 1347m, 1375s, 1551s, 3264s cm<sup>-1</sup>; its <sup>1</sup>H NMR [400MHz, CDCl<sub>3</sub>, 20°] has δ at 8.30-8.28 (m, 1H), 7.87-7.85 (m, 2H), 7.79-7.77 (m, 2H), 1.96 (s, 3H), 1.92 (s, 3H), and its <sup>13</sup>C NMR [500MHz, CDCl<sub>3</sub>, 20°] has δ at 158.9, 134.2, 133.5, 132.9, 132.0, 125.4, 25.5, 17.2. IPNSBH promotes the same reactions as NBSH except that it is more stable. A solution of 0.02M IPNSBH in DMSO is undecomposed at 50° during 30 minutes, whereas NBSH is 60% decomposed, and it is stable at 75° or 100° for 30 minutes so it can be stored at room temperature for several months without deterioration; all the same, store it under argon (as it is possibly flammable), and use gloves due to its toxicity.

It is a reagent used for allylic transposition, de-bromination and reduction of alcohols in the presence of DEAD and Ph<sub>3</sub>P. [Movassaghi et al. *Angew Chem. Int Ed Engl* 45 5859 2006, Movassaghi & Ahmad *J. Org. Chem.* 72 1838 2007].

**Isopropyl p-nitrobenzoate** [13756-40-6] **M 209.2, m 105-106°.** Dissolve it in Et<sub>2</sub>O, wash it with aqueous alkali, then H<sub>2</sub>O and dry it. Evaporate the ether and recrystallise it from EtOH to give pure material. It gives yellow crystals from petroleum ether (m 109°, 110°). [Beilstein 9 H 391, 9 II 258, 9 III 1543, 9 IV 1075.]

**Isovanillin (3-hydroxy-4-methoxybenzaldehyde)** [621-59-0] **M 152.2, m 117°, b 175°/14mm, pK<sup>25</sup> 8.89.** Crystallise isovaniline from H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>. The *oxime* has m 147°. [Beilstein 8 IV 1764.]

**Isoviolanthrone** [128-64-3] **M 456.5, m 510-511°(uncorrected).** Dissolve isoviolanthrone in 98% H<sub>2</sub>SO<sub>4</sub> and precipitate it by adding water to reduce the acid concentration to about 90%. It sublimes *in vacuo* to give dark green-violet needles [Parkyns & Ubbelohde *J Chem Soc* 4188 1960]. [Beilstein 7 I 465, 7 II 815, 7 III 4538, 7 IV 2747.]

**Janus Red B {3-[(2-hydroxy-1-naphthol)azo-2-methylphenylazo]N,N,N-trimethyl-benzenaminium chloride}** [2636-31-9] **M 460.0.** Crystallise the dye from EtOH/H<sub>2</sub>O (1:1 v/v) and dry in vacuum. Store it in a dark bottle. [Beilstein 16 II 149.]

**Ketone moschus (4-tert-butyl-2,6-dimethyl-3,5-dinitroacetophenone, Musk ketone)** [81-14-1] **M 294.3, m 134-137°, 137-138°.** Purify the ketone by recrystallisation from MeOH. It has a strong odour of musk and is used in perfumery. [Fuson et al. *J Org Chem* 12 587 1947, Beilstein 7 IV 808.]

**Lapachol [2-hydroxy-3-(3-methyl-2-butenyl)-1,4-naphthalenedione, Neutral Yellow 16]** [84-79-7] **M 242.3, CI 7549, m 140°.** Crystallise Lapachol from petroleum ether/EtOH, EtOH or Et<sub>2</sub>O. [Beilstein 8 H 326, 8 I 644, 8 II 365, 8 III 2720.]

**Leucomalachite Green** [129-73-7] **M 330.5, m 92-93°, pK<sup>25</sup> 6.90 (several pK's)**. Crystallise it from 95% EtOH (10ml/g), then from \*benzene/EtOH, and finally from petroleum ether. [Beilstein 13 H 275, 13 II 89, 13 III 135, 13 III 529, 13 481.]

**Malachite Green (carbinol)** [510-13-4] **M 346.4, m 112-114°, CI 42000, pK<sup>24</sup> 6.84**. The *oxalate* [2437-29-8] [Beilstein 8 H 326, 13 IV 2279.] is recrystallised from hot water and dried in air. The carbinol is precipitated from the oxalate (1g) in distilled water (100ml) by adding M NaOH (10ml). The precipitate is filtered off, recrystallised from 95% EtOH containing a little dissolved KOH, then washed with ether, and crystallised from petroleum ether. Dry it in a vacuum at 40°. An acid, almost colourless, solution ( $2 \times 10^{-5}$ M in  $6 \times 10^{-5}$ M H<sub>2</sub>SO<sub>4</sub>) rapidly reverts to the coloured dye. [Swain & Hedberg *J Am Chem Soc* 72 3373 1950, Beilstein 13 H 243, 744.]

**Mandelic acid (α-hydroxyphenylacetic acid)** [*S*-(+)- 17199-29-0, *R*-(-)- 611-71-2] **M 152.2, m 130-133°, 133°, 133.1° (evacuated capillary), 133-133.5°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 188° (c 5, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> (+) and (-) 155° (c 5, H<sub>2</sub>O) and (+) and (-) 158° (c 5, Me<sub>2</sub>CO), pK<sup>25</sup> 3.41**. Purify the mandelic acids by recrystallisation from H<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub>. [Roger *J Chem Soc* 2168 1932, Jamison & Turner *J Chem Soc* 611 1942.] They have solubilities in H<sub>2</sub>O of ca 11% at 25°. [Banks & Davies *J Chem Soc* 73 1938.] The *S*-benzylisothiuronium salts has **m 180° (from H<sub>2</sub>O) and [α]<sub>D</sub><sup>25</sup> (+) and (-) 57° (c 20, EtOH)** [El Masri et al. *Biochem J* 68 199 1958]. [Beilstein 10 IV 564.]

**RS-(±)-Mandelic acid** [61-72-3] **M 152.2, m 118°, 120-121°**. Purify mandelic acid by Soxhlet extraction with \*C<sub>6</sub>H<sub>6</sub> (about 6ml/g) and allow the extract to crystallise. It can also be recrystallised from CHCl<sub>3</sub>. The *S*-benzylisothiuronium salt has **m 169° (166°) (from H<sub>2</sub>O)**. Dry it at room temperature under vacuum. [Beilstein 10 IV 565.]

**Mescaline sulfate [2-(3,4,5-trimethoxyphenyl)ethylamine sulfate]** [5967-42-0] **M 309.3, m 181-184°, 183-184°, 186-189°, pK<sub>Est</sub> ~9.7**. The salt crystallises from water with 2H<sub>2</sub>O or from hot MeOH. The *acid sulfate* has **m 158°**. [Slomon & Bina *J Am Chem Soc* 68 2403 1946, Beilstein 13 I 338, 13 II 521, 13 III 2375, 13 IV 2919.]

**Mesitylene (1,3,5-trimethylbenzene)** [108-67-8] **M 120.2, m -44.7°, b 99.0-99.8°/100mm, 166.5-167°/760mm, d<sub>4</sub><sup>20</sup> 0.865, n<sub>D</sub><sup>25</sup> 1.4967**. Dry it with CaCl<sub>2</sub> and distil it from Na in a glass helices-packed column. Treat it with silica gel and redistil it. *Alternative* purifications include vapour-phase chromatography, or fractional distillation followed by azeotropic distillation with 2-methoxyethanol (which is subsequently washed out with H<sub>2</sub>O), drying and fractional distilling. More exhaustive purification uses sulfonation by dissolving in two volumes of conc H<sub>2</sub>SO<sub>4</sub>, precipitating with four volumes of conc HCl at 0°, washing with conc HCl and recrystallising from CHCl<sub>3</sub>. The mesitylene sulfonic acid is hydrolysed with boiling 20% HCl and steam distilled. The separated mesitylene is dried (MgSO<sub>4</sub> or CaSO<sub>4</sub>) and distilled. It can also be fractionally crystallised from the melt at low temperatures. [Beilstein 5 IV 1016.]

**Metanilic acid (3-aminobenzenesulfonic acid)** [121-47-1] **M 173.2, m <300°(dec), pK<sub>1</sub><sup>25</sup> <1, pK<sub>2</sub><sup>25</sup> 3.74**. Crystallise the acid from water (as the hydrate), under CO<sub>2</sub> in a semi-darkened room. The solution is photosensitive. Dry it over 90% H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. [Beilstein 14 H 688, 14 IV 2640.]

***p*-Methoxyacetophenone** [100-06-1] **M 150.2, m 39°, b 139°/15mm, 264°/736mm**. Crystallise the ketone from diethyl ether/petroleum ether. The *oxime* has **m 86°, 87° (from petroleum ether, v. Auwers et al. *Chem Ber* 58 41 1925)**. [Beilstein 8 H 87, 8 I 536, 8 II 84, 8 III 277, 8 IV 340.]

***p*-Methoxyazobenzene** [2396-60-3] **M 212.3, m 54-56°**. Crystallise 4-methoxyazobenzene from EtOH. [Beilstein 16 IV 162.]

**3-Methoxybenzanthrone** [3688-79-7] **M 274.3, m 173°**. Crystallise it from \*benzene, EtOH or Me<sub>2</sub>CO to give yellow needles. [Beilstein 8 II 239, 8 III 1629, 8 IV 1476.]

***m*-Methoxybenzoic acid (*m*-anisic acid)** [586-38-9] **M 152.2, m 110°, pK<sup>25</sup> 4.09.** Crystallise *m*-anisic acid from H<sub>2</sub>O (m 109°, 110.5°) or EtOH/water. The *S*-benzylisothiuronium salt has m 176° (from EtOH). [Beilstein 10 II 80, 10 III 244, 10 IV 316.]

***p*-Methoxybenzoic acid (*p*-anisic acid)** [100-09-4] **M 152.2, m 184.0-184.5°, pK<sup>25</sup> 4.51.** Crystallise *p*-anisic acid from EtOH, water, EtOH/water or toluene. The *S*-benzylisothiuronium salt has m 189° (from EtOH). [Beilstein 10 II 91, 10 III 280, 10 IV 346.]

**4-Methoxybenzyl chloride (anisyl chloride)** [824-94-2] **M 156.6, m -1°, b 76°/0.1mm, 95°/5mm, 110°/10mm, 117-117.5°/14mm, 117°/18mm, d<sub>4</sub><sup>20</sup> 1.15491, n<sub>D</sub><sup>20</sup> 1.55478.** Purify 4-anisyl chloride by fractional distillation under vacuum, and the middle fraction is redistilled at 10<sup>-6</sup> mm at room temperature by intermittent cooling of the receiver in liquid N<sub>2</sub>, and the middle fraction is collected. [Mohammed & Kosower *J Am Chem Soc* 93 2709 1971, Beilstein 6 IV 2137.]

**“Methoxychlor”, (1,1-bis[*p*-methoxyphenyl]-2,2,2-trichloroethane, DMDT)** [72-43-5] **M 345.7, m 78-78.2°, or 86-88°.** Free the insecticide from 1,1-bis(*p*-chlorophenyl)-2,2,2-trichloroethane by crystallising from EtOH. It is dimorphic and also crystallises from Et<sub>2</sub>O/EtOH (m 92°). [Fritsch & Feldmann *Justus Liebigs Ann Chem* 306 77 1899, Smith et al. *Aust J Chem* 29 743 1976, Beilstein 6 H 1007.]

***trans*-*p*-Methoxycinnamic acid** [830-09-1, 943-89-5 (*trans*)] **M 178.2, m 173.4-174.8°, pK<sup>25</sup> 4.54.** Crystallise the acid from MeOH to constant melting point and UV spectrum. [Beilstein 10 IV 1005.]

**6-Methoxy-1-indanone** [13623-25-1] **M 162.2, m 151-153°.** Crystallise it from MeOH, then sublime it at high vacuum. [Beilstein 8 IV 894.]

**1-Methoxy-4-nitronaphthalene** [4900-63-4] **M 203.2, m 85°.** Purify it by chromatography on silica gel, then recrystallise it from MeOH or EtOH (yellow needles). [Hodgson & Habeshaw *J Chem Soc* 47 1942, Bunce et al. *J Org Chem* 52 4214 1987, Beilstein 6 H 616, 6 III 2938.]

***p*-Methoxyphenol** [150-76-5] **M 124.1, m 54-55°, b 243°/atm, pK<sup>25</sup> 10.21.** Crystallise 4-methoxyphenol from \*benzene, petroleum ether or H<sub>2</sub>O, and dry it under vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. Sublime it *in vacuo*. [Wolfenden et al. *J Am Chem Soc* 109 463 1987, Beilstein 6 IV 5717.]

**α-Methoxyphenylacetic acid (*O*-methyl mandelic acid), [*R*(-)- 3966-32-3, *S*(+)- 26164-26-1]** **M 166.2, m 62.9°, 62-65°, 65-66°, [α]<sub>546</sub><sup>20</sup> (-) and (+) 179° (169.8°), [α]<sub>D</sub><sup>20</sup> (-) and (+) 150.7° (148°) (c 0.5, EtOH), pK<sub>Est</sub> ~3.1.** Purify the acids by recrystallising from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 80-100°). [Neilson & Peters *J Chem Soc* 1519 1962, Weizmann et al. *J Am Chem Soc* 70 1153 1948, Pirie & Smith *J Chem Soc* 338 1932, NMR: Dale & Mosher *J Am Chem Soc* 95 512 1973, for resolution: Roy & Deslongchamps *Can J Chem* 63 651 1985, Trost et al. *J Am Chem Soc* 108 4974 1986.] The *racemic mixture* has m 72°, b 121-122°/0.4mm, 165°/18mm (from petroleum ether) [Braun et al. *Chem Ber* 63 2847 1930]. [Beilstein 10 IV 566.]

***o*-Methoxyphenylacetic acid** [93-25-4] **M 166.2, m 124-125°, pK<sub>Est</sub> ~4.4.** Crystallise the acid from H<sub>2</sub>O, EtOH or aqueous EtOH, petroleum ether/Et<sub>2</sub>O and dry it in a vacuum desiccator over Sicapent. The *amide* has m 131° (from EtOH). [Beilstein 10 H 188, 10 III 422, 10 IV 536.]

***m*-Methoxyphenylacetic acid** [1798-09-0] **M 166.2, m 66-67°, 68-69°, 71.0-71.2°, pK<sub>Est</sub> ~4.3.** Crystallise the acid from H<sub>2</sub>O, or aqueous EtOH. The *S*-benzylisothiuronium salt has m 160-161° (from EtOH). [Beilstein 10 I 82, 10 III 428, 10 IV 541.]

***p*-Methoxyphenylacetic acid (homoanisic acid)** [104-01-8] **M 166.2, m 85-87°, b 138-140°/2-3mm, pK<sup>25</sup> 4.36.** Crystallise the acid from EtOH/water, EtOAc/petroleum ether (m 87°) or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (m 84-86°). [Beilstein 10 III 431, 10 IV 544.] The *acid chloride* [4693-91-8] has M 184.6, b 143°/10mm, d<sub>4</sub><sup>25</sup> 1.208. [Beilstein 10 III 434.]

*N*-(*p*-Methoxyphenyl)-*p*-phenylenediamine [101-64-4] **M 214.3, m 102°**, **b 200°/2mm, 238°/12mm, pK<sup>25</sup> 6.6 (5.9)**. Crystallise the diamine from ligroin or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (**m 99-100°**). The *picrate* has **m 164°** (from EtOH). [*Beilstein* 13 III 1161, 13 IV 1243.]

**α-Methoxy-α-trifluoromethylphenylacetic acid (MTPA, Mosher's acid)** [*R*-(+)- 20445-31-2, *S*-(-)- 17257-71-5] **M 234.2, m 43-45°, 90°/0.1mm, 105-107°/1mm, [α]<sub>D</sub><sup>20</sup> (+) and (-) 87°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 73° (c 2, MeOH), pK<sub>Est</sub> ~2.5**. A likely impurity is phenylethylamine from the resolution. Dissolve the acid in Et<sub>2</sub>O/\*benzene (3:1), wash with 0.5N H<sub>2</sub>SO<sub>4</sub>, then water, dry over magnesium sulfate, filter, evaporate and distil it. [*Dale et al. J Org Chem* 34 2543 1969, *J Am Chem Soc* 75 512 1973.]

**α-Methoxy-α-trifluoromethylphenylacetyl chloride** [*R*-(-)- 39637-99-5, *S*-(+)- 20445-33-4] **M 252.6, b 54-56°/1mm, 213-214°/760mm, d<sub>4</sub><sup>20</sup> 1.353, n<sub>D</sub><sup>20</sup> 1.468, [α]<sub>D</sub><sup>20</sup> (-) and (+) 167°, [α]<sub>D</sub><sup>20</sup> (-) and (+) 137° (c 4, CCl<sub>4</sub>), [α]<sub>D</sub><sup>24</sup> (-) and (+) 10.0° (neat)**. The most likely impurity is the free acid due to hydrolysis and should be checked by IR. If free from acid, then distil, taking care to keep moisture out of the apparatus. Otherwise add SOCl<sub>2</sub> and reflux for 5 hours and distil it. Note that shorter reflux times result in a higher boiling fraction (**b 130-155°/1mm**) which has been identified as the anhydride. [*Dale et al. J Org Chem* 34 2543 1969, for enantiomeric purity see *Dale & Mosher J Am Chem Soc* 97 512 1973.]

*N*-Methylacetanilide [579-10-2] **M 149.2, m 102-104°, b 145-146°/30mm, 225°/760mm**. Crystallise the anilide from water, ether, petroleum ether (**b 80-100°**) or \*C<sub>6</sub>H<sub>6</sub> (**m 105°**). The *BF<sub>3</sub>* salt has **m 114°** (from Me<sub>2</sub>CO/diisopropyl ether). [*Beilstein* 12 III 465, 12 IV 378.]

*p*-Methylacetophenone [122-00-9] **M 134.2, m 22-24°, b 93.5°/7mm, 110°/14mm, d<sub>4</sub><sup>20</sup> 1.000, n<sub>D</sub><sup>20</sup> 1.5335**. Impurities, including the *o*- and *m*-isomers, are removed by forming the *semicarbazone* (**m 212-213.5°**) which, after repeated crystallisation, is hydrolysed to the ketone. [*Brown & Marino J Am Chem Soc* 84 1236 1962.] It can also be purified by distillation under reduced pressure, followed by low temperature crystallisation from isopentane. The *oxime* has **m 87-88°** (from MeOH then cyclohexane *Pearson & Ball J Org Chem* 14 125 1949). [*Beilstein* 7 H 307, 7 I 154, 7 II 238, 7 III 1060, 7 IV 701.]

**1-Methylaminoanthraquinone** [82-38-2] **M 237.3, m 166.5°, pK<sub>Est</sub> ~2**. Crystallise it to constant melting point from butan-1-ol, then from EtOH. It can be sublimed under vacuum. [*Beilstein* 7 IV 2574.]

*N*-Methyl-*o*-aminobenzoic acid (*N*-methylantranilic acid) [119-68-6] **M 151.2, m 178.5°, pK<sub>1</sub><sup>25</sup> 1.97, pK<sub>2</sub><sup>25</sup> 5.34**. Crystallise the acid from water, MeOH (**m 177°**) or EtOH. [*Casulich & Smith J Am Chem Soc* 70 1923 1948, *Beilstein* 14 H 323, 14 III 895, 14 IV 1015.]

**3-(Methylamino)benzoic acid** [51524-84-6] **M 151.1, m 129°, pK<sub>1</sub><sup>25</sup> 3.08, pK<sub>2</sub><sup>25</sup> 5.10**. Crystallise the acid from H<sub>2</sub>O (needles), petroleum ether (plates) or CHCl<sub>3</sub> (**m 127°**). The *hydrochloride* has **m 244°** (plates from EtOH). [*Beilstein* 14 H 391, 14 I 559.]

*p*-Methylaminophenol sulfate (*Photol*, *Metol*) [55-55-0] **M 344.4, m 260°(dec), pK<sup>25</sup> 5.9**. Crystallise this photographic developer from H<sub>2</sub>O (needles **m 250-260°**) (its solubility is 5% at 20° and 17% at 100°) or MeOH. It is used for determining Ag. [*Palit et al. Indian J Chem, Sect B* 28 64 1989.]

*N*-Methylaniline [100-61-8] **M 107.2, b 57°/4mm, 81-82°/14mm, d<sub>4</sub><sup>20</sup> 0.985, n<sub>D</sub><sup>20</sup> 1.570, pK<sup>25</sup> 4.56**. Dry it with KOH pellets and fractionally distil it under vacuum. Acetylate, and the *acetyl* derivative is recrystallised to constant melting point (**m 101-102°**), then hydrolysed with aqueous HCl and distilled from zinc dust under reduced pressure. [*Hammond & Parks J Am Chem Soc* 77 340 1955, *Beilstein* 12 IV 241.]

*N*-Methylaniline hydrochloride [2739-12-0] **M 143.7, m 123.0-123.1°**. Crystallise the salt from dry \*benzene/CHCl<sub>3</sub> and dry under vacuum. [*Beilstein* 12 IV 241.]

**Methyl *p*-anisate** [121-98-2] **M 166.2, m 48°, b 123.8°/12mm, 24-245°/760mm**. Distil the ester and/or crystallise it from EtOH. [*Beilstein* 10 H 159, 10 III 297, 10 IV 360.]

**4-Methyl anisole** [104-93-8] **M 122.2, b 56.2°/9mm, 67°/20mm, 175-176°/760mm,  $d_{15}^{15}$  0.9757,  $n_D^{20}$  1.512.** Dissolve 4-methyl anisole in diethyl ether, wash it with M NaOH, water, dry ( $\text{Na}_2\text{CO}_3$ ), evaporate and distil it under vacuum. The *picrate* has **m 103°** (from aqueous EtOH). [Beilstein 6 IV 2098.]

**2-Methylanthracene** [613-12-7] **M 192.3, m 204-206°.** Chromatograph it on silica gel with cyclohexane as eluent and then recrystallise it from EtOH [Werst *J Am Chem Soc* 109 32 1987]. [Beilstein 5 IV 2311.]

**4-Methylanthracene** [779-02-2] **M 192.3, m 77-79°, b 196-197°/12mm,  $d_4^{20}$  1.066.** Chromatograph it on silica gel with cyclohexane as eluent and recrystallise it from EtOH [Werst *J Am Chem Soc* 109 32 1987]. [Beilstein 5 IV 2312.]

**2-Methylanthraquinone** [84-54-8] **M 222.3, m 176°, b 236-238°/10mm.** Crystallise the quinone from EtOH, then sublime it. It has  $\lambda_{\text{max}}$  at 257, 275 and 330nm (EtOH). [Hersbery & Fieser *J Am Chem Soc* 63 2562 1941, Beilstein 7 H 809, 7 III 4104, 7 IV 2574.]

**Methylarenes (see also pentamethyl- and hexamethyl-benzenes).** Recrystallise them from EtOH and sublime them in a vacuum [Schlesener et al. *J Am Chem Soc* 106 7472 1984].

**Methyl benzoate** [93-58-3] **M 136.2, b 104-105°/39mm, 199.5°/760mm,  $d_4^{20}$  1.087,  $n_D^{15}$  1.52049,  $n_D^{20}$  1.51701,  $\text{pK}^{20}$  -8.11, -6.51 ( $\text{H}_0$  scale, aqueous  $\text{H}_2\text{SO}_4$ ).** Wash the ester with dilute aqueous  $\text{NaHCO}_3$ , then water, dry with  $\text{Na}_2\text{SO}_4$  and fractionally distil it in a vacuum. [Beilstein 9 IV 283.]

***p*-Methylbenzophenone** [134-84-9] **M 196.3, m 57°, b 154-155°/3mm, 183-184°/15mm, 277-281°/3mm,  $d^{20}$  0.9926.** Crystallise the ketone from MeOH,  $\text{Et}_2\text{O}$  (m 58-59°) or petroleum ether. The *cis-oxime* has **m 154°(153-156°)** (from EtOH), and the *trans-oxime* has **m 114-116°** (from petroleum ether). [Beilstein 7 H 440, 7 III 2127, 7 IV 1403.]

**Methyl-1,4-benzoquinone (*p*-toluoquinone)** [553-97-9] **M 122.1, m 68-69°.** Crystallise *p*-toluoquinone from heptane or EtOH, dry rapidly (vacuum/ $\text{P}_2\text{O}_5$ ) and stored in a vacuum. [Beilstein 7 IV 2088.]

**Methyl benzoylformate (methyl phenylglyoxalate)** [15206-55-0] **M 164.2, m 246-248°.** Purify the ester by radial chromatography (diethyl ether/hexane, 1:1), and dry it at 110-112°/6mm. [Meyers & Oppenlaender *J Am Chem Soc* 108 1989 1986, Beilstein 10 IV 2738.]

**2-Methyl-3,4-benzphenanthrene (2-methylbenzo[*c*]phenanthrene)** [652-04-0] **M 242.3, m 70°, 80.6-81.4°, b 200°/0.4mm.** Crystallise it from EtOH (m 81-82.5°). The *picrate* has **m 118-118.5°** (yellow needles from MeOH). [Beilstein 5 III 2394, 5 IV 2570.]

***R*-(+)- $\alpha$ -Methylbenzylamine** [*R*(+) 3886-69-9, *RS*( $\pm$ ) 618-36-0] **M 121.2, b 187-188°/atm,  $[\alpha]_{546}^{20}$  +35° (c 10, EtOH),  $[\alpha]_D^{25}$  +39.7° (neat),  $\text{pK}$  9.08 (for *RS*).** Dissolve the amine in toluene, dry over NaOH and distil; fraction boiling at 187-188°/atm is collected. Store it under  $\text{N}_2$  to avoid forming the carbamate and urea. Similarly for the *S*-(-) enantiomer [2627-86-3]. [Ingersoll *Org Synth Coll Vol II* 503 1943, Robinson & Snyder *Org Synth Coll Vol III* 717 1955, Beilstein 12 IV 2424, 2425.]

***p*-Methylbenzyl chloride** [104-82-5] **M 140.6, b 80°/2mm, 98-101°/27mm,  $d_4^{20}$  1.085,  $n_D^{20}$  1.543.** Dry the chloride with  $\text{CaSO}_4$  and fractionally distil it under vacuum. [Beilstein 5 H 384, 5 III 854, 5 IV 966.]

**Methyl *o*-bromobenzoate** [610-94-6] **M 215.1, b 131.4-132°/16mm, 234-244°/760mm.** A solution of the ester in ether is washed with 10% aqueous  $\text{Na}_2\text{CO}_3$ , water, then dried and distilled. [Beilstein 9 H 348, 9 III 1385, 9 IV 1012.]

**Methyl *p*-bromobenzoate** [619-42-1] **M 215.1, m 79.5-80.5°.** Crystallise the ester from MeOH. EtOH (m 81°, also 80.5°, 79.5°) or  $\text{C}_6\text{H}_6$ /petroleum ether (m 78-79°). [Beilstein 9 H 352, 9 III 1405, 9 IV 1017.]

(±)-3'-Methyl-1,2-cyclopentenophenanthrene (16,17-dihydro-17-methyl-15*H*-cyclopenta-[*a*]-phenanthrene [549-38-2] M 232.3, m 126-127°. Crystallise it from AcOH or EtOH (plates, m 125.5-126°). The *picrate* has m 130-131° (orange-red needles from EtOH). [Tatta & Bardhan *J Chem Soc (C)* 893 1968, *Beilstein* 5 IV 2433.]

**Methyl 2,4-dichlorophenoxyacetate** [1928-38-7] M 235.1, m 43°, b 119°/11mm. Crystallise the herbicide ester from MeOH. [Branch & Jones *J Chem Soc* 2924 1955, *Beilstein* 6 III 705, 6 IV 909.]

**3,4-Methylenedioxyaniline** [14268-66-7] M 137.1, m 44.5-45.5°, 45-46°, b 108°/1mm, 144°/14mm, 156°/30mm, p*K*<sub>Est</sub> ~3.8. Crystallise the base from petroleum ether and/or distil it in a vacuum. The *hydrochloride* has m 198°(dec). [Sonn & Benirschke *Chem Ber* 54 1734 1921, *Beilstein* 19 H 328, 19 II 341, 19 III/IV 4056.]

**trans-3,4-Methylenedioxcinnamic acid** [2373-80-0] M 192.2, m 243-244°(dec), p*K*<sub>Est</sub> ~4.6. Crystallise the acid from glacial AcOH, EtOH (m 247°) or aqueous EtOH (m 240-242°), and it has m 242° after sublimation. [*Beilstein* 19 H 278, 19 II 299, 19 III/IV 3548.]

**5,5'-Methylenedisalicylic acid** [122-25-8; 27496-82-2] M 288.2, m 238°(dec). Crystallise the acid from Me<sub>2</sub>CO, \*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub>/MeOH (m 268-268°). It has λ<sub>max</sub> at 312nm (ε 7530) in EtOH. [Cushman & Kanamathareddy *Tetrahedron* 46 1406 1990.]

***N*-Methylephedrine (2-dimethylamino-1-phenylpropanol)** [1*S*,2*R*-(+)- 42151-56-4, 1*R*,2*S*-(-)- 552-79-4] M 179.3, m 85-86°, 87-87.5°, 90°, b 115°/2mm, [α]<sub>D</sub><sup>20</sup> (+) and (-) 35°, [α]<sub>D</sub><sup>20</sup> (+) and (-) 30° (c 4.5, MeOH), p*K*<sub>26</sub> 9.22. *N*-Methylephedrine has been recrystallised from Et<sub>2</sub>O, petroleum ether, of aqueous EtOH or aqueous MeOH and has been distilled under reduced pressure. [Smith *J Chem Soc* 2056 1927, Tanaka & Sugawa *Yakugaku Zasshi (J Pharm Soc Japan)* 72 1548 1952 (*Chem Abstr* 47 8682 1953), Takamatsu *Yakugaku Zasshi (J Pharm Soc Japan)* 76 1227 1956, *Chem Abstr* 51 4304 1957.] The *hydrochloride* has m 192-193° and [α]<sub>D</sub><sup>20</sup>+30° (c 5, H<sub>2</sub>O) [Prelog & Hüfliger *Helv Chim Acta* 33 2021 1950]. [*Beilstein* 13 IV 1884.]

**Methyl gallate** [99-24-1] M 184.2, m 202°. Crystallise the gallate ester from MeOH. [*Beilstein* 10 IV 1998.]

**Methyl Green** [82-94-0, 7114-03-6 (*ZnCl*<sub>2</sub> salt)] M 458.5, m >200°(dec). Crystallise the dye from hot water. [*Beilstein* 13 IV 2286.]

**Methyl 4-hydroxybenzoate** [99-76-3] M 152.2, m 127.5°, p*K*<sub>Est</sub> ~9.3. Fractionally crystallise the ester from its melt, and recrystallise it from \*benzene, then from \*benzene/MeOH and dry it over CaCl<sub>2</sub> in a vacuum desiccator. [*Beilstein* 10 IV 360.]

**Methyl 3-hydroxy-2-naphthoate** [883-99-8] M 202.2, m 73-74°, p*K*<sub>Est</sub> ~9.0. Crystallise the ester from MeOH (charcoal) containing a little water. [*Beilstein* 10 IV 1186.]

**3-Methylmercaptoaniline** [1783-81-9] M 139.2, b 101.5-102.5°/0.3mm, 163-165°/16mm, d<sub>4</sub><sup>20</sup> 1.147, n<sub>D</sub><sup>20</sup> 1.641, p*K*<sub>25</sub> 4.05. Purify the aniline by fractional distillation in an inert atmosphere. It has UV max at 226 and 300nm. [Bordwell & Cooper *J Am Chem Soc* 74 1058 1952.] The *N*-*acetyl* derivative has m 78-78.5° (from aqueous EtOH). The *hydrochloride* has m 260-261° (aqueous EtOH/HCl) or m 225-227° (EtOH/Et<sub>2</sub>O). [*Beilstein* 13 H 533, 13 III 1221, 13 IV 1289.]

**4-Methylmercaptoaniline** [104-96-1] M 139.2, b 140°/15mm, 151°/25mm, 155°/23mm, d<sub>4</sub><sup>20</sup> 1.137, n<sub>D</sub><sup>20</sup> 1.639, p*K*<sub>25</sub> 4.40. Purify the aniline by fractional distillation in an inert atmosphere. The *hydrochloride* has m 242-246° (from aqueous EtOH/HCl). The *sulfone* has m 137° (from H<sub>2</sub>O), p*K*<sub>25</sub> 1.48, and the *sulfone hydrochloride* has m 260-261° (from aqueous EtOH/HCl). [Lumbroso & Passerini *Bull Soc Chim Fr* 311 1957, Mangini & Passerini *J Chem Soc* 4954 1956, *Beilstein* 13 H 533, 13 II 297, 13 IV 1221.]

**1-Methylnaphthalene** [90-12-0] M 142.2, f -30°, b 244.6°, d<sub>4</sub><sup>20</sup> 1.021, n<sub>D</sub><sup>20</sup> 1.6108. Dry 1-methylnaphthalene for several days with CaCl<sub>2</sub> or by prolonged refluxing with BaO. Fractionally distil it through a glass helices-

packed column from sodium. Purify it further by solution in MeOH and precipitation of its *picrate complex* by adding to a saturated solution of picric acid in MeOH. The picrate, after crystallisation to constant melting point (**m** 140-141°) from MeOH, is dissolved in \*benzene and extracted with aqueous 10% LiOH until the extract is colourless. Evaporation of the \*benzene solution under vacuum gives 1-methylnaphthalene [Kloetzel & Herzog *J Am Chem Soc* **72** 1991 1950]. However, neither the picrate nor the styphnate complexes satisfactorily separate 1- and 2-methylnaphthalenes. To achieve this, 2-methylnaphthalene (10.7g) in 95% EtOH (50ml) has been precipitated with 1,3,5-trinitrobenzene (7.8g) and this complex has been crystallised from MeOH to **m** 153-153.5° (**m** of the 2-methyl isomer is 124°). [Alternatively, 2,4,7-trinitrofluorenone in hot glacial acetic acid could be used, and the derivative (**m** 163-164°) is recrystallised from glacial acetic acid]. The 1-methylnaphthalene is regenerated by passing a solution of the complex in dry \*benzene through a 15-in column of activated alumina and washing with \*benzene/petroleum ether (b 35-60°) until the coloured band of the nitro compound had moved down near the end of the column. The complex can also be decomposed using tin and acetic-hydrochloric acids, followed by extraction with diethyl ether and \*benzene; the extracts are washed successively with dilute HCl, strongly alkaline sodium hypophosphite, water, dilute HCl and water. [Soffer & Stewart *J Am Chem Soc* **74** 567 1952.] It can be freed from anthracene by zone melting [Beilstein **5** IV 1687.]

**2-Methylnaphthalene** [91-57-6] **M 142.2, m 34.7-34.9°, b 129-130°/25mm.** Fractionally crystallise repeatedly from its melt, then fractionally distil under reduced pressure. It has been crystallised from \*benzene and dried under vacuum in an Abderhalden pistol. It can be purified *via* its *picrate* (**m** 114-115°) or better *via* the 1,3,5-trinitrobenzene complex as for 1-methylnaphthalene (above). [Beilstein **5** IV 1693.]

**6-Methyl-2-naphthol** [17579-79-2] **M 158.2, m 128-129°, b 177.5-178°/15mm, pK<sub>Est</sub> ~9.8.** Crystallise the naphthol from EtOH or ligroin. Sublime it *in vacuo*. [Beilstein **6** II 618, **6** III 3028.]

**7-Methyl-2-naphthol** [26593-50-0] **M 158.2, m 118°, pK<sub>Est</sub> ~9.7.** Crystallise the naphthol from EtOH or ligroin. It has **m** 118° after sublimation *in vacuo*. [Halsall & Thomas *J Chem Soc* 2564 1956, Beilstein **6** IV 3029.]

**Methyl 1-naphthyl ether (1-methoxynaphthalene)** [2216-69-5] **M 158.2, b 90-91°/2mm, d<sub>4</sub><sup>20</sup> 1.095, n<sub>D</sub><sup>26</sup> 1.6210.** Steam distil the ether from alkaline solution. The distillate is extracted with Et<sub>2</sub>O. After drying (MgSO<sub>4</sub>) the extract and evaporating Et<sub>2</sub>O, the methyl naphthyl ether is then fractionated under reduced pressure from CaH<sub>2</sub>. The *picrate* has **m** 129.5-130.5° (from EtOH). [Beilstein **6** IV 4211.]

**Methyl 2-naphthyl ether (2-methoxynaphthalene, Nerolin)** [93-04-9] **M 158.2, m 73.0-73.6°, b 138°/10mm 273°/760mm.** Fractionally distil the ether under vacuum. Crystallise it from absolute EtOH, aqueous EtOH, \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or *n*-heptane, and dry it under vacuum in an Abderhalden pistol or distil it *in vacuo*. The *picrate* has **m** 118° (from EtOH or CHCl<sub>3</sub>). [Kikuchi et al. *J Phys Chem* **91** 574 1987, Beilstein **6** III 2969, **6** IV 4257.]

**2-Methyl-3-nitroaniline** [603-83-8] **M 152.2 m 92°, b 305°/760mm, pK<sub>Est</sub> ~2.3.** Crystallise the nitrotoluidine from EtOH or \*C<sub>6</sub>H<sub>6</sub>. It is steam volatile. The *acetyl* derivative crystallises from aqueous EtOH and has **m** 164°. [Beilstein **12** I 395, **12** II 460, **12** III 1944, **2** IV 1811.]

**N-Methyl-4-nitroaniline** [100-15-2] **M 152.2, m 152.2°, pK<sub>25</sub> 0.55.** Crystallise the aniline from aqueous EtOH. [Beilstein **12** H 714.]

**2-Methyl-4-nitroaniline** [99-52-5] **M 152.2 m 129°, pK<sub>25</sub> 0.93.** Crystallise the nitrotoluidine from EtOH. The *acetyl* and *benzoyl* derivatives have **m** 200° and 174° (EtOH) respectively. [Beilstein **12** IV 1809.]

**2-Methyl-5-nitroaniline** [99-55-8] **M 152.2, m 109°, pK<sub>25</sub> 2.35.** Acetylate the aniline, and the acetyl derivative is crystallised to constant melting point; then hydrolyse it with 70% H<sub>2</sub>SO<sub>4</sub> and the free base is regenerated by treatment with NH<sub>3</sub> [Bevan et al. *J Chem Soc* 4284 1956]. [Beilstein **12** H 844, **12** IV 1807.]

**4-Methyl-3-nitroaniline** [119-32-4] **M 152.2, m 81.5°, pK<sub>25</sub> 3.02.** Crystallise the aniline from hot water

(charcoal), then ethanol and dry it in a vacuum desiccator. [*Beilstein* 12 H 966.]

**2-Methyl-5-nitroanisole** [4837-88-1] **M 176.2, m 54-56°**. 2-Methyl-5-nitroanisole crystallises from MeOH (yellow needles) and sublimes *in vacuo*. [Kuffner *Monatsh Chem* 91 1152 1960, *Beilstein* 6 I 178.]

**Methyl 3-nitrobenzoate** [618-95-1] **M 181.2, m 78°**. Crystallise the benzoate from MeOH (1g/ml). [*Beilstein* 9 H 378, 9 I 153, 9 II 248, 9 III 1493, 9 IV 1056.]

**Methyl 4-nitrobenzoate** [619-50-1] **M 181.2, m 95-95.5°**. Dissolve the benzoate in diethyl ether, then wash it with aqueous alkali; the ether is evaporated and the ester is recrystallised from EtOH. [*Beilstein* 9 H 390, 9 IV 1074.]

**3-Methyl-2-nitrobenzoic acid** [5437-38-7] **M 181.2, m 220-222.5°, pK<sup>20</sup> 2.91 (1% aqueous EtOH)**. Recrystallise it from EtOH. The *methyl ester* has **m 74°** (from MeOH), and the *amide* [60310-07-8] **M 180.1**, has **m 192°** (needles from H<sub>2</sub>O, prisms from EtOH). [*Beilstein* 9 H 480, 9 IV 1722.]

**4-Methyl-3-nitrobenzoic acid (3-nitro-*p*-toluic acid)** [96-98-0] **M 181.2, m 190-191°, pK<sup>20</sup> 3.62 (1% aqueous EtOH)**. Recrystallise the acid from EtOH. The *S*-benzylisothiuronium salt has **m 167-168°** (EtOH). The *acid chloride* [10397-30-5] has **m 20-21°, b 185°/36mm**, and the *methyl ester* [7356-11-8] crystallises as pale yellow needles from MeOH with **m 51°**. [*Beilstein* 9 H 502, 9 II 334, 9 III 2359.]

***N*-Methyl-4-nitrosoaniline** [10595-51-4] **M 136.2, m 114-115°, 118°, pK<sub>Est</sub> ~1.0**. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>. The *picrate* has **m 166°(dec)** (from MeOH or CHCl<sub>3</sub>). [*Beilstein* 7 III 3370, 12 IV 1228.]

***N*-Methyl-*N*-nitroso-*p*-toluenesulfonamide (Diazald)** [80-11-5] **M 214.2, m 62°**. Crystallise diazald from \*benzene by addition of petroleum ether and store it in a refrigerator. It is soluble in most organic solvents including Et<sub>2</sub>O, and liberates diazomethane on treatment with alkali. Store it in the cold. It remains active for many months if it is kept in a stoppered bottle in a cold room at >10°. [deBoer & Backer *Org Synth* 34 96 1954, *Org Synth Coll Vol VI* 943 1963, *Beilstein* 11 I 29.]

**4-Methylphenylacetic acid (*p*-tolylacetic acid)** [622-47-9] **M 150.2, m 94°, pK<sup>25</sup> 4.37**. Crystallise the acid from heptane or water. [*Beilstein* 9 IV 1795.]

**1-Methyl-1-phenylhydrazine sulfate** [33008-18-3] **M 218.2, pK<sup>25</sup> 4.98 (free base)**. Crystallise the sulfate from hot H<sub>2</sub>O by addition of hot EtOH. [*Beilstein* 15 IV 53 for free base.]

***N*-Methylphthalimide** [550-44-7] **M 161.1, m 133.8°**. Recrystallise the imide from absolute EtOH or AcOH (**m 134°**). The IR has  $\nu_{\max}$  at 1780 and 1380cm<sup>-1</sup>. [*Beilstein* 21 H 461, 21 III/IV 5030.]

**Methyl Red (4-dimethylaminoazobenzene-2'-carboxylic acid)** [493-52-7] **M 269.3, m 181-182°, CI 13020, pK<sub>1</sub><sup>25</sup> 2.30, pK<sub>2</sub><sup>25</sup> 4.82**. The acid is extracted with boiling toluene using a Soxhlet apparatus. The crystals which separate on slow cooling to room temperature are filtered off, washed with a little toluene and recrystallised from glacial acetic acid, \*benzene or toluene followed by pyridine/water. *Alternatively*, dissolve it in aqueous 5% NaHCO<sub>3</sub> solution, and precipitate it from a hot solution by dropwise addition of aqueous HCl. Repeat this until the extinction coefficients do not increase. [*Beilstein* 16 IV 504.]

**Methyl salicylate (methyl 2-hydroxybenzoate, oil of winter green)** [119-36-8] **M 152.2, m -8.6°, b 79°/6mm, 104-105°/14mm, 223.3°/atm, d<sub>4</sub><sup>20</sup> 1.1149, n<sub>D</sub><sup>20</sup> 1.5380, pK<sup>25</sup> 10.19**. Dilute the ester with Et<sub>2</sub>O, wash with saturated NaHCO<sub>3</sub> (it may effervesce due to the presence of free acid), brine, dry MgSO<sub>4</sub>, filter, evaporate and distil it. Its solubility is 1g/1.5L of H<sub>2</sub>O. The *benzoyl* derivative has **m 92° (b 270-280°/120mm)**, and the *3,5-dinitrobenzoate* has **m 107.5°**, and the *3,5-dinitrocarbamoyl* derivative has **m 180-181°**. [Hallas *J Chem Soc* 5770 1965, *Beilstein* 10 IV 143.]

**$\alpha$ -Methylstyrene (monomer, 2-phenylpropene)** [98-83-9] **M 118.2, b 57°/15mm, d<sub>4</sub><sup>20</sup> 0.910, n<sub>D</sub><sup>20</sup> 1.5368**.

Wash the monomer three times with aqueous 10% NaOH (to remove inhibitors such as quinol), then six times with distilled water, dry with  $\text{CaCl}_2$  and distil it under vacuum. The distillate is kept under nitrogen, in the cold, and redistilled if kept for more than 48 hours before use. It can also be dried with  $\text{CaH}_2$ . Add stabiliser if it is to be stored for long periods. [Beilstein 5 IV 1364.]

**trans- $\beta$ -Methylstyrene (1-phenylpropene)** [873-66-5] M 118.2, b 64-65°/10mm, 176°/760mm,  $d_4^{20}$  0.910,  $n_D^{20}$  1.5496. Distil it under  $\text{N}_2$  from powdered NaOH through a Vigreux column, and pass it through activated neutral alumina before use [Wong et al. *J Am Chem Soc* 109 3428 1987]. [Beilstein 5 III 1184, 5 IV 1359.]

**4-Methylstyrene** [622-97-9] M 118.2, b 60°/12mm, 106°/10mm,  $d_4^{20}$  0.9173,  $n_D^{20}$  1.542. Purify it as the above styrenes and add a small amount of antioxidant if it is to be stored. It has UV in EtOH at  $\lambda_{\text{max}}$  at 285nm (log  $\epsilon$  3.07), and in EtOH + HCl 295nm (log  $\epsilon$  2.84) and 252nm (log  $\epsilon$  4.23). [Schwartzman & Carson *J Am Chem Soc* 78 322 1956, Joy & Orchin *J Am Chem Soc* 81 305 1959, Buck et al. *J Chem Soc* 2377 1949, Beilstein 5 IV 1369.]

**Methyl 4-toluenesulfonate** [80-48-8] M 186.2, m 25-28°, 28°, b 144.6-145.2°/5mm, 168-170°/13mm,  $d_4^{20}$  1.23,  $n_D^{20}$  1.5172. The ester is purified by distillation *in vacuo* and could be crystallised from petroleum ether or  $\text{Et}_2\text{O}$ /petroleum ether at low temperature. It is a powerful methylating agent, is **TOXIC** and is a **skin irritant**, so it is better to purify it by repeated distillation. [IR: Schreiber *Anal Chem* 21 1168 1949, Buehler et al. *J Org Chem* 2 167 1937, Roos et al. *Org Synth Coll Vol I* 145 1941, Beilstein 11 IV 247.]

**Methyl Violet 2B [4,4'-bis-(diethylamino)-4''-methyliminotriphenylmethyl hydrochloride]** [8004-87-3] M 394.0, m 137°(dec), CI 42535,  $\lambda_{\text{max}}$  ~580nm. Crystallise the dye from EtOH by precipitation with  $\text{Et}_2\text{O}$  during cooling in an ice-bath. Filter it off and dry it at 105°. [Beilstein 13 IV 2283.]

**Michler's ketone [4,4'-bis(dimethylamino)benzophenone]** [90-94-8] M 268.4, m 179°,  $\text{pK}^{25}$  9.84. Dissolve the ketone in dilute HCl, filter and precipitate it by adding ammonia (to remove water-insoluble impurities such as benzophenone). Then crystallise it from EtOH or petroleum ether. [Suppan *J Chem Soc, Faraday Trans I* 71 539 1975.] It is also purified by dissolving in \*benzene, then washing with water until the aqueous phase is colourless. The \*benzene is evaporated off, and the residue is recrystallised three times from \*benzene and EtOH [Hoshino & Kogure *J Phys Chem* 72 417 1988]. [Beilstein 14 IV 255.]

**Naphthacene (benz[b]anthracene, 2,3-benzanthracene, rubene)** [92-24-0] M 228.3, m >300°, 341° (open capillary), 349°, 357°. Naphthacene crystallises in orange needles from EtOH,  $^*\text{C}_6\text{H}_6$  or toluene. Dissolve it in sodium-dried \*benzene and pass it through a column of alumina. The eluent is evaporated under vacuum, and the chromatography is repeated using fresh \*benzene. Finally, the naphthacene is sublimed *in vacuo* at 186°. [Martin & Ubbelohde *J Chem Soc* 4948 1961, UV: Clar *Chem Ber* 65 517 1932, Clar *Chem Ber* 69 607 1936, IR: Cannon & Sutherland *Spectrochim Acta* 4 373 1951, Beilstein 5 H 718, 5 IV 2545.]

**1-Naphthaldehyde** [66-77-3] M 156.2, m 2°, b 160-161°/15mm,  $\text{pK}^{20}$  -7.04 (aqueous  $\text{H}_2\text{SO}_4$ ). Distil 1-naphthaldehyde with steam, extract the distillate into  $\text{Et}_2\text{O}$ , dry ( $\text{Na}_2\text{SO}_4$ ), filter, evaporate the filtrate and distil the residue in a vacuum. [Beilstein 7 IV 1286.]

**2-Naphthaldehyde** [66-99-9] M 156.2, m 59°, b 260°/19mm,  $\text{pK}^{20}$  -7.04 (aqueous  $\text{H}_2\text{SO}_4$ ). Distil 2-naphthaldehyde with steam then crystallise it from water or EtOH. [Beilstein 7 IV 1288.]

**Naphthalene** [91-20-3] M 128.2, m 80.3°, b 87.5°/10mm, 218.0°/atm,  $d_4^{20}$  1.0253,  $d_4^{100}$  0.9625,  $n_D^{85}$  1.5590. Crystallise naphthalene once or more times from the following solvents: EtOH, MeOH,  $\text{CCl}_4$ ,  $^*\text{C}_6\text{H}_6$ , glacial acetic acid, acetone or diethyl ether, followed by drying at 60° in an Abderhalden drying apparatus. It has also been purified by vacuum sublimation and by fractional crystallisation from its melt. Other purification procedures include refluxing in EtOH over Raney Ni and chromatography of a  $\text{CCl}_4$  solution on alumina with \*benzene as eluting solvent. Baly and Tuck [*J Chem Soc* 1902 1908] purified naphthalene for spectroscopy by

heating with conc  $\text{H}_2\text{SO}_4$  and  $\text{MnO}_2$ , followed by steam distillation (repeating the process), and formation of the *picrate* which, after recrystallisation ( $m$   $150^\circ$ ) is decomposed with base and the naphthalene is steam distilled. It is then crystallised from dilute EtOH. It can be dried over  $\text{P}_2\text{O}_5$  under vacuum (take care not to make it sublime). Also purify it by sublimation and subsequent crystallisation from cyclohexane. *Alternatively*, it has been washed at  $85^\circ$  with 10% NaOH to remove phenols, with 50% NaOH to remove nitriles, with 10%  $\text{H}_2\text{SO}_4$  to remove organic bases, and with 0.8g  $\text{AlCl}_3$  to remove thianaphthalenes and various alkyl derivatives. Then it is treated with 20%  $\text{H}_2\text{SO}_4$ , 15%  $\text{Na}_2\text{CO}_3$  and finally distilled. [Gorman et al. *J Am Chem Soc* **107** 4404 1985.] Zone refining purified naphthalene from anthracene, 2,4-dinitrophenylhydrazine, methyl violet, benzoic acid, methyl red, chrysene, pentacene and indoline. [*Beilstein* **5** IV 1640.]

**Naphthalene-1,4-disulfonic acid** [92-41-1] **M 288.2**,  $\text{pK}_{\text{Est}} < 0$ . Crystallise the acid from conc HCl. The *disulfonamide* has  $m$   $273^\circ$  (from EtOH). [*Beilstein* **11** II 119, **11** III 463.]

**Naphthalene-1-sulfonic acid** [85-47-2] **M 208.2**,  $m$  ( $2\text{H}_2\text{O}$ )  $90^\circ$ , (anhydrous)  $139\text{--}140^\circ$ ,  $\text{pK}^{20} -0.17$ . Crystallise the acid from conc HCl and twice from  $\text{H}_2\text{O}$ . The *S-benzylisothiuronium salt* has  $m$   $137^\circ$  (from aqueous EtOH). [*Beilstein* **11** H 155, **11** III 383, **11** IV 521.]

**Naphthalene-2-sulfonic acid** [120-18-3] **M 208.2**,  $m$   $91^\circ$ ,  $\text{pK}_{\text{Est}} < 1$ . Crystallise the acid from conc HCl. The *S-benzylisothiuronium salt* has  $m$   $192^\circ$  (from aqueous EtOH). [Berger *Acta Chem Scand* **8** 432 1954, *Beilstein* **11** H 171, **11** IV 527.]

**Naphthalene-1-sulfonyl chloride** [85-46-1] **M 226.7**,  $m$   $64\text{--}67^\circ$ ,  $68^\circ$ ,  $b$   $147.5^\circ/0.9\text{mm}$ ,  $147.5^\circ/13\text{mm}$ . If the IR indicates the presence of OH, then treat it with an equal weight of  $\text{PCl}_5$  and heat it at  $ca$   $100^\circ$  for 2 hours, cool and pour onto ice +  $\text{H}_2\text{O}$ , stir well and filter off the solid. Wash the solid with cold  $\text{H}_2\text{O}$  and dry the solid in a vacuum desiccator over  $\text{P}_2\text{O}_5$  + solid KOH. Extract the solid with petroleum ether ( $b$   $40\text{--}60^\circ$ ), filter off any insoluble solid and cool. Collect the crystalline sulfonyl chloride and recrystallise it from dry  $\text{Et}_2\text{O}$ , petroleum ether or  $^*\text{C}_6\text{H}_6$ /petroleum ether. If large quantities are available, then it can be distilled under high vacuum. [Fierz-Davaid & Weissenbach *Helv Chim Acta* **3** 2312 1920.] The *sulfonamide* crystallises from EtOH ( $m$   $150.5^\circ$ ) or  $\text{H}_2\text{O}$  ( $m$   $153^\circ$ ). [*Beilstein* **11** H 175, **11** II 93, **11** IV 383.]

**Naphthalene-2-sulfonyl chloride** [93-11-8] **M 226.7**,  $m$   $74\text{--}76^\circ$ ,  $78^\circ$ ,  $79^\circ$ ,  $83^\circ$ ,  $b$   $148^\circ/0.6\text{mm}$ ,  $201^\circ/13\text{mm}$ . Distil the chloride in a vacuum and/or recrystallise it (twice) from  $^*\text{benzene}$ /petroleum ether (1:1 v/v). Purify it as the 1-sulfonyl chloride above. [Fierz-Davaid & Weissenbach *Helv Chim Acta* **3** 2312 1920.] The *sulfonamide* has  $m$   $217^\circ$  (from EtOH). [*Beilstein* **11** III 399, **11** IV 529.]

**1,8-Naphthalic acid (naphthalene-1,8-dicarboxylic acid)** [518-05-8] **M 216.9**,  $m$   $270^\circ$ ,  $\text{pK}_{\text{Est}(1)} \sim 2.1$ ,  $\text{pK}_{\text{Est}(2)} \sim 4.5$ . Crystallise the acid from EtOH or aqueous EtOH. [Raecke & Schrip *Org Synth* **40** 71 1960, *Beilstein* **9** II 651, **9** III 4466.]

**1,8-Naphthalic anhydride** [81-84-5] **M 198.2**,  $m$   $274^\circ$ ,  $274\text{--}275^\circ$ . Extract it with cold aqueous  $\text{Na}_2\text{CO}_3$  to remove free acid, then crystallise from acetic anhydride. [*Beilstein* **17** III/IV 6392, **17/11** V 492.]

**2-Naphthamide (2-naphthoic acid amide)** [2243-82-5] **M 171.2**,  $m$   $195^\circ$ ,  $\text{pK}^{20} -2.30$  ( $\text{H}_0$  scale, aqueous  $\text{H}_2\text{SO}_4$ ). Crystallise it from EtOH ( $197^\circ$ ). [Clemo & Spence *J Chem Soc* 2818 1928, *Beilstein* **9** H 657, **9** II 45, **9** IV 2417.]

**Naphthazarin (5,8-dihydroxy-1,4-naphthoquinone)** [475-38-7] **M 190.2**,  $m \sim 220\text{--}230^\circ$  (dec),  $m$   $225\text{--}230^\circ$ ,  $\text{pK}_1^{20} 8.2$ ,  $\text{pK}_2^{20} 10.2$ . It crystallises in red-brown needles with a green shine from EtOH. It also crystallises from hexane and is further purified by sublimation at 2-10mm. [Huppert et al. *J Phys Chem* **89** 5811 1985.] It is sparingly soluble in  $\text{H}_2\text{O}$  but soluble in alkalis. The *diacetate* forms golden yellow prisms from  $\text{CHCl}_3$ ,  $m$   $192\text{--}193^\circ$  and the *5,8-dimethoxy* derivative has  $m$   $157^\circ$  ( $155^\circ$ ) (from petroleum ether) [Bruce & Thompson *J Chem Soc* 1089 1955, IR: Schmand & Boldt *J Am Chem Soc* **97** 447 1975, NMR: Brockmann & Zeeck *Chem Ber* **101** 4221 1968]. The *monothiosemicarbazone* has  $m$   $168^\circ$  (dec) from EtOH [Gardner et al. *J Am Chem Soc* **74** 2106 1952]. [*Beilstein* **8** H 412, **8** III 3600.]

**Naphthionic acid (4-aminonaphthalene-1-sulfonic acid)** [84-86-6] **M 223.3, m > 300<sup>o</sup>(dec), pK<sup>25</sup> 2.68.** It crystallises from H<sub>2</sub>O as needles of the 0.5 hydrate. Salt solutions fluoresce strongly blue. The *S*-benzylisothiuronium salt has **m 195<sup>o</sup>** (from aqueous EtOH). [Beilstein 14 IV 2793.]

**1-Naphthoic acid** [86-55-5] **M 172.2, m 162.5-163.0<sup>o</sup>, pK<sup>25</sup> 3.60.** Crystallise the acid from toluene (3ml/g) (charcoal), petroleum ether (b 80-100<sup>o</sup>), or aqueous 50% EtOH. The *amide* has **m 202<sup>o</sup>** (from EtOH). [Beilstein 9 IV 2402.]

**2-Naphthoic acid** [93-09-4] **M 172.2, m 184-185<sup>o</sup>, pK<sup>25</sup> 4.14.** Crystallise the acid from EtOH (4ml/g), aqueous 50% EtOH or Me<sub>2</sub>CO (**m 185-186<sup>o</sup>**). Dry it at 100<sup>o</sup>. The *acid chloride* has **m 52-52<sup>o</sup>** (from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether) and **b 160-162<sup>o</sup>/11mm** [Hersberg & Carson *Org Synth Coll Vol III* 629 1955], the *amide* has **m 192<sup>o</sup>** (from EtOH), and the *N-Methyl amide* has **m 109-109.5<sup>o</sup>** (form \*C<sub>6</sub>H<sub>6</sub>). [Beilstein 9 H 656, 9 III 3174, 9 IV 2402.]

**1-Naphthol** [90-15-3] **M 144.2, m 95.5-96<sup>o</sup>, pK<sup>25</sup> 9.34.** Sublime 1-naphthol, then crystallise it from aqueous MeOH (charcoal), aqueous 25% or 50% EtOH, \*C<sub>6</sub>H<sub>6</sub>, cyclohexane, heptane, CCl<sub>4</sub> or H<sub>2</sub>O. Dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. The *4-nitrobenzoate* has **m 143<sup>o</sup>** (from EtOH). [Shizuka et al. *J Am Chem Soc* 107 7816 1985, Beilstein 8 H 596, 6 IV 4208.]

**2-Naphthol** [135-19-3] **M 144.2, m 122.5-123.5<sup>o</sup>, pK<sup>25</sup> 9.57.** Crystallise 2-naphthol from aqueous 25% EtOH (charcoal), H<sub>2</sub>O, \*benzene, toluene or CCl<sub>4</sub>. *Alternatively*, extract it repeatedly with small amounts of EtOH, followed by dissolution in a minimum volume of EtOH and precipitation with distilled water, then drying over P<sub>2</sub>O<sub>5</sub> under vacuum. It has also been dissolved in aqueous NaOH and precipitated by adding acid (repeat several times), then precipitated from \*benzene by addition of heptane. Final purification can be by zone melting or sublimation *in vacuo*. The *4-nitrobenzoate* has **m 104<sup>o</sup>** (from EtOH). [Bardez et al. *J Phys Chem* 89 5031 1985, Kikuchi et al. *J Phys Chem* 91 574 1987, Beilstein 6 IV 4253.]

**Naphthol AS-D (3-hydroxy-2-naphthoic-*o*-toluide)** [135-61-5] **M 277.3, m 196-198<sup>o</sup>.** Purify it by recrystallisation from xylene. This gives yellow-green fluorescent solutions at pH 8.2-9.5. [IR: Schnopper et al. *Anal Chem* 31 1542 1959.] The *naphthol AS-D acetate* is obtained with AcCl, **m 168-169<sup>o</sup>**, and with chloroacetyl chloride *naphthol AS-D-chloroacetate* is obtained [Moloney et al. *J Histochem Cytochem* 8 200 1960, Burstone *Arch Pathology* 63 164 1957]. [Beilstein 12 H 505.]

**α-Naphtholbenzein [bis-(α-{4-hydroxynaphth-1-yl})-benzyl alcohol]** [6948-88-5] **M 392.5, m 122-125<sup>o</sup>, pK<sub>Est</sub> ~ 9.3.** Crystallise the alcohol from EtOH, aqueous EtOH or glacial acetic acid. [Beilstein 6 H 1150.]

**1-Naphthol-2-carboxylic acid (1-hydroxy-2-naphthoic acid)** [86-48-6] **M 188.2, m 203-204<sup>o</sup>, pK<sub>Est(1)</sub> ~ 2.5, pK<sub>Est(2)</sub> ~ 12.** Successively crystallise the acid from EtOH/water, diethyl ether and acetonitrile, with filtration through a column of charcoal and Celite. [Tong & Glesmann *J Am Chem Soc* 79 583 1957, Beilstein 10 H 331, 10 IV 1194.]

**3-Naphthol-2-carboxylic acid (3-hydroxy-2-naphthoic acid)** [92-70-6] **M 188.2, m 222-223<sup>o</sup>, pK<sub>1</sub><sup>25</sup> 2.79, pK<sub>2</sub><sup>25</sup> 12.84.** Crystallise it from water or acetic acid. The *S*-benzylisothiuronium salt has **m 216-217<sup>o</sup>** (from EtOH). It forms many metal complex salts. [Beilstein 10 H 333, 10 III 1084.]

**1,2-Naphthoquinone** [524-42-5] **M 158.2, m 140-142<sup>o</sup>(dec).** Crystallise the quinone from ether (red needles) or \*benzene (orange leaflets). [Beilstein 7 IV 2417.]

**1,4-Naphthoquinone** [130-15-4] **M 158.2, m 125-125.5<sup>o</sup>.** Crystallise the quinone from diethyl ether (charcoal). It distils in steam. It also crystallises from \*benzene or aqueous EtOH and sublimes in a vacuum. [Beilstein 7 IV 2422.]

**β-Naphthoxyacetic acid** [120-23-0] **M 202.2, m 156<sup>o</sup>, pK<sub>Est</sub> ~ 3.0.** Crystallise the acid from hot water or \*benzene. [Beilstein 6 IV 4274.]

**$\beta$ -Naphthoyltrifluoroacetone (4,4,4-trifluoro-2-naphthylbutan-1,3-dione)** [893-33-4] **M 266.2, m 70-71 $^{\circ}$ , 74-76 $^{\circ}$ , pK $^{20}$  6.35.** Crystallise the dione from EtOH. The *mono oxime* crystallises from H $_2$ O or aqueous EtOH and has **m 137-138 $^{\circ}$ .** [Reid & Calvin *J Am Chem Soc* **72** 2948 1950, *Beilstein* **7** IV 2478.]

**Naphthvalene (2,3-dihydro-1,2,3-metheno-1H-indene)** [34305-47-0] **M 128.1, m dec at 175 $^{\circ}$  to benzvalene.** Purify it by chromatography on alumina and eluting with pentane, also by reverse phase (ODS) HPLC using MeCN as solvent. It is stable at room temperature [Kjell et al. *J Am Chem Soc* **108** 4111 1986, Abelt et al. *J Am Chem Soc* **107** 4148 1985]. The  $^1\text{H}$  NMR in CCl $_4$  has  $\tau$  3.18 (4H), 6.17 (t  $J$  1.5Hz, 2H), 7.60 (t  $J$  1.5Hz 2H).

**1-Naphthyl acetate** [830-81-9] **M 186.2, m 45-46 $^{\circ}$ , 48-49 $^{\circ}$ , b 114-115 $^{\circ}$ /1mm.** Chromatograph the acetate on silica gel and crystallise as the 2-isomer below. The *2,4,7-trinitrofluoren-9-one complex* has **m 120 $^{\circ}$**  (from EtOH). [*Beilstein* **6** H 608, **6** III 2928, **6** IV 4217.]

**2-Naphthyl acetate** [1523-11-1] **M 186.2, m 71 $^{\circ}$ , 71-72 $^{\circ}$ , b 133-134 $^{\circ}$ /2mm.** Distil the acetate in a high vacuum and/or crystallise it from petroleum ether (b 60-80 $^{\circ}$ ) or dilute aqueous EtOH. The *picrate* has **m 80 $^{\circ}$**  (from EtOH). [*Beilstein* **6** H 644, **6** I 313, **6** II 600, **5** III 2982, **6** IV 4267.]

**1-Naphthylacetic acid** [86-87-3] **M 186.2, m 132 $^{\circ}$ , pK $^{25}$  4.23.** Crystallise the acid from EtOH or water. [*Beilstein* **9** H 666, **9** IV 2424.]

**2-Naphthylacetic acid** [581-96-4] **M 186.2, m 143.1-143.4 $^{\circ}$ , pK $^{25}$  4.30.** Crystallise the acid from water or \*benzene. [*Beilstein* **9** H 666, **9** IV 2431.]

**1-Naphthylamine** [134-32-7] **M 143.2, m 50.8-51.2 $^{\circ}$ , b 160 $^{\circ}$ , pK $^{25}$  3.94.** Sublime the amine at 120 $^{\circ}$  in a stream of nitrogen, then crystallise it from petroleum ether (b 60-80 $^{\circ}$ ), or absolute EtOH then diethyl ether. Dry it *in vacuo* in an Abderhalden pistol. It has also been purified by crystallisation of its *hydrochloride* (see below) from water, followed by liberation of the free base and distillation; it is finally purified by zone melting. The *styphnate* has **m 181-182 $^{\circ}$**  (from EtOH). [*Beilstein* **12** III 2846, **12** IV 3009.] **CARCINOGEN.**

**1-Naphthylamine hydrochloride** [552-46-5] **M 179.7, m 240-250 $^{\circ}$  sublimes on heating.** Crystallise the salt from water (charcoal). [*Beilstein* **12** III 2849, **12** IV 3009.]

**2-Naphthylamine** [91-59-8] **M 143.2, m 113 $^{\circ}$ , pK $^{25}$  4.20.** Sublime the amine at 180 $^{\circ}$  in a stream of nitrogen. Crystallise it from hot water (charcoal) or \*benzene. Dry it under vacuum in a drying pistol. The *styphnate* has **m 194-195 $^{\circ}$**  (from EtOH). [*Beilstein* **12** H 1265, **12** III 2989, **12** IV 3122.] **CARCINOGEN.**

**2-Naphthylamine-1-sulfonic acid** [81-16-3] **M 223.3, m >200 $^{\circ}$ (dec), pK $_1^{25}$  <1, pK $_2^{25}$  2.35 (NH $_2$ ).** Crystallise the acid under nitrogen from boiling water and dry it in a steam oven [Bryson *Trans Faraday Soc* **47** 522, 527 1951]. [*Beilstein* **14** III 2240, **14** IV 2792.]

**5-Naphthylamine-1-sulfonic acid** [84-89-9] **M 223.3, m >200 $^{\circ}$ (dec), pK $_{\text{Est}(1)} <1$ , pK $_2^{25}$  3.69 (NH $_2$ ).** Crystallise the acid under nitrogen from boiling water and dry it in a steam oven [Bryson *Trans Faraday Soc* **47** 522, 527 1951]. [*Beilstein* **14** IV 2800.]

**2-Naphthylamine-6-sulfonic acid** [93-00-5] **M 223.3, m >200 $^{\circ}$ (dec), pK $^{25}$  3.74.** Crystallise the acid from a large volume of hot water. The *diethylamine salt* has **m 190.5-192 $^{\circ}$**  (from EtOH/*iso*BuOH), and the *S*-*benzylisothiuronium salt* has **m 330 $^{\circ}$**  (from *n*BuOH). [*Beilstein* **14** H 760, **14** II 463, **14** III 2249, **14** IV 2804.]

**1-(1-Naphthyl) ethanol** [*R*-(+)- 42177-25-3, *S*-(-)- 15914-84-8] **M 172.2, m 46 $^{\circ}$ , 45-47.5 $^{\circ}$ , 48 $^{\circ}$ , [ $\alpha$ ] $_{546}^{20}$  (+) and (-) 94 $^{\circ}$ , [ $\alpha$ ] $_D^{20}$  (+) and (-) 78 $^{\circ}$  (c 1, MeOH).** Purify the alcohol by recrystallisation from Et $_2$ O/petroleum ether, Et $_2$ O, hexane [Balfe et al. *J Chem Soc* 797 1946, IR, NMR: Theisen & Heathcock *J Org Chem* **53** 2374 1988, see also Fredga et al. *Acta Chem Scand* **11** 1609 1957]. The *RS-alcohol* [57605-95-5] has **m 63-65 $^{\circ}$ , 65-66 $^{\circ}$**  from hexane. [*Beilstein* **6** III 3034, **6** IV 4346.]

**1-(1-Naphthyl)ethylamine** [*R*-(+)- 3886-70-2, *S*-(-)- 10420-89-0] **M 171.2, b 153°/11mm, 178-181°/20mm,  $d_4^{20}$  1.067,  $n_D^{20}$  1.624,  $[\alpha]_{546}^{20}$  (+) and (-) 65°,  $[\alpha]_D^{20}$  (+) and (-) 55° (c 2, MeOH),  $[\alpha]_D^{17}$  (+) and (-) 82.8° (neat),  $pK_{Est} \sim 9.3$ .** Purify the amine by distillation in a good vacuum. [Mori et al. *Tetrahedron* **37** 1343 1981, cf Wilson in *Topics Stereochem* (Allinger and Eliel eds) **vol 6** 135 1971, Fredga et al. *Acta Chem Scand* **11** 1609 1957.] The *hydrochlorides* crystallise from H<sub>2</sub>O [ $\alpha$ ]<sub>D</sub><sup>18</sup> ±3.9° (c 3, H<sub>2</sub>O), and the *sulfates* recrystallise from H<sub>2</sub>O as *tetrahydrates* **m 230-232°**. The *RS-amine* [42882-31-5] has **b 153°/11mm, 156°/15mm, 183.5°/41mm** [Blicke & Maxwell *J Am Chem Soc* **61** 1780 1939]. [*Beilstein* **12** III 3111.]

***N*-( $\alpha$ -Naphthyl)ethylenediamine dihydrochloride** [1465-25-4] **M 291.2, m 188-190°,  $pK_{Est(1)} \sim 3.8$ ,  $pK_{Est(2)} \sim 9.4$ .** Crystallise the salt from water. [*Beilstein* **12** II 699.]

**1-Naphthyl isocyanate** [86-84-0] **M 169.2, m 3-5°, b 140-142°/12mm, 269-270°/760mm,  $d_4^{20}$  1.1774.** Distil the isocyanate at atmospheric pressure or in a vacuum. It can be crystallised from petroleum ether (b 60-70°) at low temperature. *It has a pungent odour, is TOXIC and is absorbed through the skin.* [*Beilstein* **12** H 1244, **12** III 2948.]

**1-Naphthyl isothiocyanate** [551-06-4] **M 185.3, m 58-59°.** Crystallise the isothiocyanate from hexane (1g in 9 ml). It forms white needles and is soluble in most organic solvents but is insoluble in H<sub>2</sub>O. *It is absorbed through the skin and may cause dermatitis.* [Cymmerman-Craig et al. *Org Synth Coll Vol IV* 700 1963, *Beilstein* **12** H 1244, **12** III 2948.]

**2-(2-Naphthoxy)ethanol** [93-20-9] **M 188.2, m 72-74°, 76.7°.** Crystallise it from \*benzene/petroleum ether. [Yoshino et al. *Bull Chem Soc Jpn* **46** 553 1973.]

***N*-1-Naphthylphthalamic acid (Naptalam)** [132-66-1] **M 291.3, m 189° (dec), 203°.** Crystallise the herbicide from EtOH (**m 183-185°**). The *Na salt* has **m 185°**. [*Beilstein* **12** H 1236, **12** I 525, **12** III 2876.]

**2-Naphthyl salicylate (Betol)** [613-78-5] **M 264.3, m 94.4°, 95°,  $pK_{Est} \sim 10.0$ .** Crystallise Betol from EtOH. It is *dimorphic* with **m**'s at  $\sim 55^\circ$  and  $\sim 96^\circ$ . [*Beilstein* **10** H 80, **10** II 53, **10** III 136, **10** IV 158.]

**1-Naphthyl thiourea (ANTU)** [86-88-4] **M 202.2, m 198°.** Crystallise ANTU from EtOH. [*Beilstein* **12** III 2941, **12** IV 3086.]

**1-Naphthyl urea** [6950-84-1] **M 186.2, m 215-220°.** Crystallise the urea from EtOH (**m 213-214°** also 215°). [*Beilstein* **12** H 1238, **12** IV 3076.]

**2-Naphthyl urea** [13114-62-0] **M 186.2, m 212°, 219-220°.** Crystallise the urea from EtOH. [*Beilstein* **12** H 1292, **12** III 3029, **12** IV 3149.]

**Narcein** {6-[6-(2-dimethylaminoethyl)-2-methoxy-3,4-(methylenedioxy)phenylacetyl]-2,3-dimethoxybenzoic acid} [131-28-2] **M 445.4, m 176-177° (145° anhydrous),  $pK_1^{15}$  3.5,  $pK_2^{15}$  9.3.** Recrystallise Narcein from water (as trihydrate). The *stypnate* has **m 185-189°** (from EtOH), and the *picrate* has **m 200°** (from EtOH). [*Beilstein* **19** H 370, **19** I 797, **19** II 386, **19** IV 4382.]

**Neostigmine [(3-dimethylcarbamoylphenyl)trimethylammonium] bromide** [114-80-7] **M 303.2, m 176°(dec), 181°(dec).** Crystallise neostigmine bromide from EtOH/diethyl ether. Its solubility in H<sub>2</sub>O is  $\sim 50\%$ . [*Beilstein* **13** III 939.] (It is cholinergic and *highly TOXIC*.) The starting material **3-dimethylcarbamoyl-*N,N*-dimethylaniline** [59-99-4] has **b 195°/20mm** [*Beilstein* **13** III 936], and its *picrate* has **m 138°** (from EtOH).

**Neostigmine methyl sulfate (Prostigmine B)** [51-60-5] **M 334.4, m 142-145°.** Crystallise the sulfate from EtOH or Me<sub>2</sub>CO (**m 143-144°**). Its solubility in H<sub>2</sub>O is  $\sim 10\%$ . [*Beilstein* **13** III 939.] (It is cholinergic and *highly TOXIC*.)

**Ninhydrin (1,2,3-triketohydrindene hydrate)** [485-47-2] **M 178.1, m 241-243°(dec), pK<sup>30</sup> 8.82.** Crystallise ninhydrin from hot water (charcoal). Dry it under vacuum and store it in sealed brown containers. [Beilstein 7 IV 2786.]

**2-Nitroacetanilide** [552-32-9] **M 180.2, m 93-94°, pK<sub>Est</sub> <0.** Crystallise the anilide from H<sub>2</sub>O, aqueous EtOH (m 92-93°) or EtOH (m 92.5-93.5°). [Beilstein 12 II 371, 12 III 1523, 12 IV 1574.]

**4-Nitroacetanilide** [104-04-1] **M 180.2, m 217°, pK<sub>Est</sub> <0.** Precipitate the anilide from 80% H<sub>2</sub>SO<sub>4</sub> by adding ice, then wash with water, and crystallise from aqueous EtOH. Dry it in air. [Beilstein 12 IV 1632.]

**3-Nitroacetophenone** [121-89-1] **M 165.2, m 81°, b 167°/18mm, 202°/760mm.** Distil the ketone in steam and crystallise it from EtOH. [Beilstein 7 IV 656.]

**4-Nitroacetophenone** [100-19-6] **M 165.2, m 80-81°, b 145-152°/760mm.** Crystallise the ketone from EtOH or aqueous EtOH. [Beilstein 7 IV 657.]

**3-Nitroalizarin (1,2-dihydroxy-3-nitro-9,10-anthraquinone, Alizarin Orange)** [568-93-4] **M 285.2, m 244°(dec), pK<sub>Est(1)</sub> ~4.6, pK<sub>Est(2)</sub> ~9.6.** Crystallise the dye from AcOH (m 244-245°). It has  $\lambda_{\text{max}}$  at 494nm (H<sub>2</sub>SO<sub>4</sub>) and forms Cu salts. [Beilstein 8 H 447, 8 II 491, 8 III 3774, 8 IV 2359.]

***o*-Nitroaniline** [88-74-4] **M 138.1, m 72.5-73.0°, pK<sup>25</sup> -0.25 (-0.31).** Crystallise the aniline from hot water (charcoal), then from aqueous 50% EtOH, or EtOH, and dry it in a vacuum desiccator. It has also been chromatographed on alumina, then recrystallised from \*benzene. [Beilstein 12 IV 1563.]

***m*-Nitroaniline** [99-09-2] **M 138.1, m 114°, pK<sup>25</sup> 2.46.** Purify it as for *o*-nitroaniline. **Warning: it is absorbed through the skin.** [Beilstein 12 IV 1589.]

***p*-Nitroaniline** [100-01-6] **M 138.1, m 148-148.5°, pK<sup>25</sup> 1.02.** Purify it as for *o*-nitroaniline. It also crystallises from acetone. It is freed from *o*- and *m*-isomers by zone melting and sublimation. [Beilstein 12 IV 1613.]

***o*-Nitroanisole (2-methoxynitrobenzene)** [91-23-6] **M 153.1, f 9.4°, b 265°/737mm, d<sub>4</sub><sup>20</sup> 1.251, n<sub>D</sub><sup>20</sup> 1.563.** Purify it by repeated vacuum distillation in the absence of oxygen. [Beilstein 6 IV 1249.]

***p*-Nitroanisole (4-methoxynitrobenzene)** [100-17-4] **M 153.1, m 54°.** Crystallise it from petroleum ether or hexane and dry it *in vacuo*. [Beilstein 6 IV 1282.]

**9-Nitroanthracene** [602-60-8] **M 223.2, m 142-143°.** Purify it by recrystallisation from EtOH or MeOH. Further purify it also by sublimation or TLC. [Beilstein 5 H 666, 5 II 578, 5 III 2136, 5 IV 2296.]

***o*-Nitrobenzaldehyde** [552-89-6] **M 151.1, m 44-45°, b 120-144°/3-6mm.** Crystallise the aldehyde from toluene (2-2.5ml/g) by addition of 7ml petroleum ether (b 40-60°) for 1ml of solution. It can also be distilled under reduced pressures. [Beilstein 7 IV 584.]

***m*-Nitrobenzaldehyde** [99-61-6] **M 151.1, m 58°.** Crystallise the aldehyde from water or EtOH/water, then sublime it twice at 2mm pressure at a temperature slightly above its melting point. [Beilstein 7 IV 591.]

***p*-Nitrobenzaldehyde** [555-16-8] **M 151.1, m 106°.** Purify it as for *m*-nitrobenzaldehyde above. [Beilstein 7 IV 589.]

**Nitrobenzene** [98-95-3] **M 123.1, f 5.8°, b 84-86.5°/6.5-8mm, 210.8°/760mm, d<sub>4</sub><sup>20</sup> 1.206, n<sub>D</sub><sup>15</sup> 1.55457, n<sub>D</sub><sup>20</sup> 1.55257, pK<sup>18</sup> -11.26 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Common impurities include nitrotoluene, dinitrothiophene, dinitrobenzene and aniline. Most impurities can be removed by steam distillation in the presence of dilute H<sub>2</sub>SO<sub>4</sub>, followed by drying with CaCl<sub>2</sub>, and shaking with, then distilling at low pressure from BaO, P<sub>2</sub>O<sub>5</sub>, AlCl<sub>3</sub>

or activated alumina. It can also be purified by fractional crystallisation from absolute EtOH (by refrigeration). Another purification process includes extraction with aqueous 2M NaOH, then water, dilute HCl, and water, followed by drying (CaCl<sub>2</sub>, MgSO<sub>4</sub> or CaSO<sub>4</sub>) and fractional distillation under reduced pressure. The pure material is stored in a brown bottle, in contact with silica gel or CaH<sub>2</sub>. It is very *hygroscopic*. [Beilstein 5 H 233, 5 I 124, 5 II 171, 5 III 591, 5 IV 708.]

**4-Nitrobenzene-azo-resorcinol (magneson-I)** [74-39-5] M 259.2, m 185°(dec), 199-200°(dec). Crystallise the dye from EtOH. [Beilstein 16 H 181, 16 IV 1266.]

**4-Nitrobenzenediazonium fluoroborate** [456-27-9] M 236.9. It crystallises from water. Drain it well and use it as soon as possible. Do NOT store it. [Whetsch et al. *J Am Chem Soc* 78 3360 1956, Beilstein 16 IV 816.] **It can be EXPLOSIVE when dry.**

**2-Nitrobenzenesulfonyl chloride (NPS-Cl)** [7669-54-7] M 189.6, m 73-74.5°, 74.5-75°, 74-76°. Recrystallise it from CCl<sub>4</sub> (2ml/g), filter off the solution at 5° (recovery 75%). It has also been recrystallised from petroleum ether (b 40-60°), dried rapidly at 50° and stored in a brown glass bottle, sealed well and stored away from moisture. [Hubacher *Org Synth Coll Vol II* 455 1943, Ito et al. *Chem Pharm Bull Jpn* 26 296 1978, Beilstein 6 I 157.]

***o*-Nitrobenzenesulfonyl chloride** [1694-92-4] M 221.6, m 63-67°, 68-69°, 69.5-69.7°. It is prepared by the oxidation of di-*o*-nitrophenyldisulfide with Cl<sub>2</sub> [Wertheim *Org Synth Coll Vol II* 471 1943] or from *o*-nitrobenzenesulfonic acid and PCl<sub>5</sub>, then pouring into H<sub>2</sub>O, washing the solid with H<sub>2</sub>O, drying it *in vacuo* and recrystallising it from petrol, Et<sub>2</sub>O/petrol, petroleum ether/CCl<sub>4</sub> or petroleum ether. With ammonia it is converted to the **sulfonamide** [5455-59-4] M 202.1, m 193°, which forms needles from hot aqueous EtOH; and with phenol/pyridine or Na<sub>2</sub>CO<sub>3</sub>, the **phenyl ester**, m 57°, is obtained which crystallises from EtOH. The **sulfonylazide**, m 71-73°, separates as needles from petroleum ether [Leffler et al. *J Org Chem* 28 902 1963]. [Beilstein 11 H 67, 11 I 20, 11 II 32, 11 III 114, 11 IV 174.]

***m*-Nitrobenzenesulfonyl chloride** [121-51-7] M 221.6, m 61-62°, 63.9-64.2°. It can be prepared from *m*-nitrobenzenesulfonic acid and PCl<sub>5</sub> as for the isomer above and crystallised from petroleum ether/CCl<sub>4</sub> or petroleum ether in needles. With ammonia it is converted to the **sulfonamide** [121-52-8] M 202.1, m 167-168°, which forms needles from hot aqueous EtOH and is a moderately strong base with **pK<sup>20</sup> 9.20**. With phenol/pyridine or Na<sub>2</sub>CO<sub>3</sub>, the **phenyl ester**, m 94°, is obtained which crystallises from acetic acid. [Gilbert *Synthesis* 315 1977, Beilstein 11 H 69, 11 I 21, 11 II 33, 11 III 126, 11 IV 182.]

***p*-Nitrobenzenesulfonyl chloride (Nosyl chloride)** [98-74-8] M 221.6, m 66-70°, 78-80°, 80°, b 143-144°/1.5mm, 180-181°/17mm. It is prepared in the same way as for the *m*-isomer above from *p*-nitrobenzenesulfonic acid, and is purified by vacuum distillation and/or by recrystallisation from petroleum ether or Et<sub>2</sub>O/petroleum ether. With ammonia it is converted to the **sulfonamide** [6325-93-5] M 202.1, m 179-180°, which forms prisms from hot aqueous EtOH and is a moderately strong base with **pK<sup>20</sup> 9.14**. [Prinsen et al. *Rec Trav Chim Pays Bas* 84 24 1965, Beilstein 11 H 72, 11 I 21, 11 II 34, 11 III 136, 11 IV 192.]

It is highly efficient in solid phase peptide synthesis when using *N-nosyl- $\alpha$ -amino acids* [Leggio et al. *Tetrahedron* 63 8164 2007.]

***o*-Nitrobenzenesulfonylhydrazine (NBSH)** [606-26-8] M 217.2, m 100-101°, 101° (dec). NBSH is unstable unless handled properly. If the material is suspect, it is preferable to prepare freshly thus: hydrazine monohydrate (12.1ml, 2.5 equivalents) is added to a solution of *o*-nitrobenzenesulfonyl chloride (22.2g, 1 equivalent) in dry THF (100ml) at -30° under argon whereby the solution became brown, white N<sub>2</sub>H<sub>4</sub>.HCl deposited and after 30 minutes TLC (2/1, EtOAc/hexanes) indicated that all the sulfonyl chloride was used up. EtOAc (200ml, at 23°) is added to the cold mixture and washed repeatedly with ice-cold 10% aqueous NaCl (5 × 150ml, contact time with each wash should not be >1 minute; the use of just distilled H<sub>2</sub>O results in much lower yields). The organic layer is washed, dried (Na<sub>2</sub>SO<sub>4</sub> at 0°), then added slowly to a stirred solution of hexanes (1.2L) at 23° during 5 minutes. The off-white NBSH that precipitates within 10 minutes is isolated by vacuum filtration, the solid is washed with hexanes (2 × 50ml, 23°) and dried at 23°/ <1.5mm for 14 hours to provide pure

NBSH as an off-white powder (17.6g, 81%; TLC has  $R_F$  0.19, EtOAc/hexanes 2/1). Store it at  $-0-5^\circ$  under argon, and use gloves due to its toxicity. Its IR (EtOAc) has  $\nu_{\max}$  at 1165, 1352, 1547, 3400—3100  $\text{cm}^{-1}$  and its  $^1\text{H}$  NMR [300MHz,  $\text{CD}_3\text{CN}$ ] has  $\delta$ : 8.17-8.03 (1H, m), 7.91-7.78 (3H, m), 5.97 (1H, bs), 3.90 (2H, bs); and its  $^{13}\text{C}$  NMR [75MHz,  $\text{CD}_3\text{CN}$ ] has  $\delta$ : 149.4, 135.5, 133.4, 133.2, 130.8 and 125.8. [Meyers et al. *J Org Chem* **62** 7507 1997, Dunn & Davies *J Chem Soc* 1050 1929.]

NBSH is a versatile reagent for synthesising allenes from propargyl alcohols [Meyers & Zheng *J Am Chem Soc* **118** 4492 1996], reductive transposition of allylic alcohols [Meyers & Zheng *Tetrahedron Lett* **37** 4841 1996], and deoxygenation of unhindered alcohols; each of which proceeds by Mitsunobu displacement of an alcohol with NBSH followed by *in situ* elimination of *o*-nitrobenzoylsulfinic acid to produce mono-alkyl diazene intermediates under mild reaction conditions, i.e. neutral pH and less than  $23^\circ$ , particularly useful with substrates that have sensitive substituents. [Mitsunobu *Synthesis* 1 1981, Hughes *Org React* **42** 335 1992.] For the success of these reactions it is imperative that the reagent is as pure as possible. Solutions of NBSH in THF that are pale yellow are still suitable for these applications.

**4-Nitrobenzhydrazide** [606-26-8] **M 181.1, m 124 $^\circ$** . Crystallise the hydrazide from EtOH or EtOAc (**m 120 $^\circ$** ). [Beilstein **9** H 375, **9** I 152, **9** II 246, **9** III 1481, **9** IV 1052.]

**2-Nitrobenzoic acid** [552-16-9] **M 167.1, m 146-148 $^\circ$ , pK $^{25}$  2.21**. Crystallise the acid from \*benzene (twice), *n*-butyl ether (twice), then water (twice). Dry and store it in a vacuum desiccator. [Le Noble & Wheland *J Am Chem Soc* **80** 5397 1958.] It has also been crystallised from EtOH/ $\text{H}_2\text{O}$ . The *amide* has **m 176.5 $^\circ$**  (from  $\text{H}_2\text{O}$ ). [Beilstein **9** III 1466, **9** IV 1046.]

**3-Nitrobenzoic acid** [121-92-6] **M 167.1, m 143-143.5 $^\circ$ , pK $^{25}$  3.46**. Crystallise the acid from \*benzene,  $\text{H}_2\text{O}$ , EtOH (charcoal), glacial acetic acid or MeOH/ $\text{H}_2\text{O}$ . Dry and store it in a vacuum desiccator. The *amide* has **m 143 $^\circ$**  (from  $\text{H}_2\text{O}$  or \* $\text{C}_6\text{H}_6$ ). [Beilstein **9** III 1489, **9** IV 1055.]

**4-Nitrobenzoic acid** [62-23-7] **M 167.1, m 241-242 $^\circ$ , pK $^{25}$  3.43**. Purify it as for 3-nitrobenzoic acid above. The *amide* has **m 201.6 $^\circ$**  (from  $\text{H}_2\text{O}$ ). [Beilstein **9** III 1537, **9** IV 1072.]

**4-Nitrobenzoyl chloride** [122-04-3] **M 185.6, m 75 $^\circ$ , b 155 $^\circ$ /20mm**. Crystallise the acid chloride from dry petroleum ether (b 60-80 $^\circ$ ), \* $\text{C}_6\text{H}_6$  or  $\text{CCl}_4$ . Distil it under vacuum. **Irritant**. [Adama & Jenkins *Org Synth Coll Vol I* 394 1941, Beilstein **9** III 1709, **9** IV 1191.]

**4-Nitrobenzyl alcohol** [619-73-8] **M 153.1, m 93 $^\circ$** . Crystallise the alcohol from EtOH and sublime it *in vacuo*. Purity should be at least 99.5%. Sublimed samples should be stored in the dark over anhydrous  $\text{CaSO}_4$  (Drierite). If the IR contains OH bands, then the sample should be resublimed before use. [Mohammed & Kosower *J Am Chem Soc* **93** 2709 1979, Beilstein **6** IV 2611.]

**4-Nitrobenzyl bromide** [100-11-8] **M 216.0, m 98.5-99.0 $^\circ$** . Recrystallise the bromide four times from absolute EtOH, then twice from cyclohexane/hexane/\*benzene (1:1:1), followed by sublimation at 0.1mm and final recrystallisation from the same solvent mixture. [Lichtin & Rao *J Am Chem Soc* **83** 2417 1961.] It has also been crystallised from petroleum ether (b 80-100 $^\circ$ , 10ml/g, charcoal). It slowly decomposes even when stored in a desiccator in the dark. **IRRITANT**. [Beilstein **5** IV 861.]

***m*-Nitrobenzyl chloride** [619-23-8] **M 171.6, m 45 $^\circ$** . Crystallise the chloride from petroleum ether (b 90-120 $^\circ$ ). **IRRITANT**. [Beilstein **5** IV 855.]

***p*-Nitrobenzyl chloride** [100-14-1] **M 171.6, m 72.5-73 $^\circ$** . Crystallise the chloride from  $\text{CCl}_4$ , dry diethyl ether, or *n*-heptane, and dry it under vacuum. **IRRITANT**. [Beilstein **5** IV 856.]

***p*-Nitrobenzyl cyanide** [555-21-5] **M 162.2, m 117 $^\circ$** . Crystallise the nitrile from EtOH. **TOXIC**. [Beilstein **9** H 456, **9** I 183, **9** II 313, **9** III 2291.]

**2-Nitrobiphenyl** [86-00-0] **M 199.2, m 36.7 $^\circ$ , 37 $^\circ$ , b 185-195 $^\circ$ /20-30mm**. *o*-Nitrobiphenyl has been prepared

by a Gomberg reaction between diazotised 2-nitroaniline and  $*C_6H_6$  [Elks et al. *J Chem Soc* 1282 1940] or by the nitration of biphenyl [Morgan & Walls *J Soc Chem Ind* 49 15T 1930]. Although the respective yields are ~45% and ~25%, the latter method is by far more convenient to perform experimentally. Crystallise it from EtOH (seeding required) or petroleum ether (b 40-60°). Sublime it under vacuum. [*Beilstein* 5 H 582, 5 I 273, 5 II 487, 5 III 1750.]

**3-Nitrocinnamic acid** [555-68-0] **M 193.2, m 200-201°, pK<sup>25</sup> 2.58 (trans)**. Crystallise the acid from  $*C_6H_6$  or EtOH. The *p*-bromophenacyl ester has **m 178°** (from AcOH). [*Beilstein* 9 III 271, 9 IV 2043.]

**4-Nitrocinnamic acid** [882-06-4] **M 193.2, m 143° (cis), 286°(trans), pK<sub>Est</sub> ~2.6 (trans)**. Crystallise the acid from H<sub>2</sub>O. The *p*-bromophenacyl ester has **m 191°** (from AcOH). [*Beilstein* 9 H 606, 9 III 2744.]

**4-Nitrodiphenylamine** [836-30-6] **M 214.2, m 133-134°, pK<sup>25</sup> -2.5**. Crystallise the amine from EtOH or aqueous EtOH (**m 135-136°**) and has **m 131.5-132°** after sublimation *in vacuo*. [*Beilstein* 12 H 715, 12 III 1586.]

**2-Nitrodiphenyl ether** [2216-12-8] **M 215.2, b 106-108°/0.01mm, 137-138°/0.5mm, 161-162°/4mm, 188-189°/12mm, 195-200°/25mm, d<sub>4</sub><sup>20</sup> 1.241, n<sub>D</sub><sup>25</sup> 1.600**. Purify the ether by fractional distillation. Its UV (EtOH) has  $\lambda_{max}$  at: 255, 315nm ( $\epsilon$  6200 and 2800), and the IR (CS<sub>2</sub>) has bands at 1350 (NO<sub>2</sub>) and 1245, 1265 (COC) cm<sup>-1</sup> [UV, IR: Dahlgard & Brewster *J Am Chem Soc* 80 5861 1958, Tomita & Takase *Yakugaku Zasshi (J Pharm Soc Japan)* 75 1077 1955, Fox & Turner *J Chem Soc* 1115 1930, Henley *J Chem Soc* 1222 1930]. [*Beilstein* 6 H 218, 6 II 210, 6 III 801.]

**Nitrodurene (1,2,4,5-tetramethyl-3-nitrobenzene)** [3463-36-3] **M 179.2, m 113-114°, b 143-144°/10mm**. Distil nitrodurene in a vacuum and/or crystallise it from EtOH (yellow prisms, **m 113-113.5°**), MeOH, acetic acid, petroleum ether or chloroform. It has been crystallised by dissolving in hexane and kept at -20° for crystals to form (**m 111-112°**). Its UV has  $\lambda_{max}$  at 238 and 400nm (iso-octane). [Masnovi et al. *J Am Chem Soc* 111 2263 1989, cf Powell & Johnson *Org Synth Coll Vol II* 449 1943, *Beilstein* 5 H 432, 5 III 982, 5 IV 1080.]

**3-Nitrofluoranthene** [892-21-7] **M 247.3, m 159-160°**. Recrystallise it from AcOH or EtOAc (yellow crystals) and/or sublime it at high vacuum. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>, Me<sub>2</sub>CO or  $*C_6H_6$ , soluble in warm EtOH and very soluble in Et<sub>2</sub>O. [Kloetzel et al. *J Am Chem Soc* 78 1165 1956, *Beilstein* 5 III 2279.]

**2-Nitrofluorene** [607-57-8] **M 211.2, m 156°**. Crystallise 2-nitrofluorene from aqueous acetic acid or Me<sub>2</sub>CO (**m 158-158.5°**, also 160-160.5°). [*Beilstein* 5 H 628, 5 III 1948.]

**Nitromesitylene (2-nitro-1,3,5-trimethylbenzene)** [603-71-4] **M 165.2, m 44°, b 255°/760mm**. Crystallise it from EtOH, or a small volume of MeOH and cool in an ice-salt bath (**m 43.5°**). [Powell & Johnson *Org Synth Coll Vol II* 449 1943, *Beilstein* 5 H 410, 5 III 923, 5 IV 1028.]

**Nitromethylphenylsulfone** [21272-85-5] **M 201.2, m 78-79°**. Recrystallise the sulfone from CHCl<sub>3</sub>, or 95% EtOH, then sublime it at 100-120°/0.1 Torr and recrystallise again. It is an acidic analogue of nitromethane used for the general synthesis of aliphatic acids and nitriles and in cycloaddition reactions with olefins to form heterocycles. [Wade et al. *J Org Chem* 46 765 1981, 49 4595 1984, *Beilstein* 6 II 292.]

**1-Nitronaphthalene** [86-57-7] **M 173.2, m 57.3-58.3°, b 30-40°/0.01mm**. Fractionally distil 1-nitronaphthalene under reduced pressure, then crystallise it from EtOH, aqueous EtOH or heptane. Chromatograph it on alumina with  $*benzene$ /petroleum ether as eluent. It sublimes *in vacuo*. The 1:1 picric acid complex has **m 72°** (from EtOH). [*Beilstein* 5 H 553, 5 III 1593, 5 IV 1673.]

**2-Nitronaphthalene** [581-89-5] **M 173.2, m 79°, b 165°/15mm**. Distil it in a vacuum and/or crystallise it from aqueous EtOH and sublime in a vacuum. The 1:1 1,3,5-trinitrobenzene complex has **m 75.5°** (from EtOH). [*Beilstein* 5 H 555, 5 III 1596, 5 IV 1675.]

**1-Nitro-2-naphthol** [550-60-7] **M 189.2, m 103°**, **b 115°/0.05mm, pK<sup>25</sup> 5.93**. Distil it under high vacuum and/or crystallise the naphthol (repeatedly) from \*benzene/petroleum ether (b 60-80°)(1:1). [*Beilstein* 6 H 653, 6 III 3002, 6 IV 4370.]

**2-Nitro-1-naphthol** [607-24-9] **M 189.2, m 127-128°**, **pK<sup>25</sup> 5.89**. Crystallise the naphthol (repeatedly) from EtOH. [*Beilstein* 6 H 615, 6 III 2938, 6 IV 4236.]

**2-Nitrophenol** [88-75-5] **M 139.1, m 44.5-45.5°**, **b 214-215°/760mm, pK<sup>25</sup> 7.23**. Crystallise 2-nitrophenol from EtOH/water, water, EtOH, \*benzene or MeOH/petroleum ether (b 70-90°). It can be steam distilled. Petrucci and Weygandt [*Anal Chem* 33 275 1961] crystallised it from hot water (twice), then EtOH (twice), followed by fractional crystallisation from the melt (twice), drying over CaCl<sub>2</sub> in a vacuum desiccator and then in a drying pistol. The *4-nitrobenzoate* had **m 141°** (from EtOH). [*Beilstein* 6 IV 1246.]

**3-Nitrophenol** [554-84-7] **M 139.1, m 96°**, **b 160-165°/12mm, pK<sup>25</sup> 8.36**. Crystallise 3-nitrophenol from water, CHCl<sub>3</sub>, CS<sub>2</sub>, EtOH or petroleum ether (b 80-100°), and dry it under vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. It can also be distilled at low pressure. The *4-nitrobenzoate* had **m 174°** (from EtOH). [*Beilstein* 6 IV 1269.]

**4-Nitrophenol** [100-02-7] **M 139.1, m 113-114°**, **pK<sup>25</sup> 7.16**. Crystallise 4-nitrophenol from water (which may be acidified, e.g. with *N* H<sub>2</sub>SO<sub>4</sub> or 0.5*N* HCl), EtOH, aqueous MeOH, CHCl<sub>3</sub>, \*benzene or petroleum ether, then dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub> at 25°. It can be sublimed at 60°/10<sup>-4</sup>mm. The *4-nitrobenzoate* had **m 159°** (from EtOH). [*Beilstein* 6 IV 1279.]

**2-Nitrophenoxyacetic acid** [1878-87-1] **M 197.2, m 150-159°**, **158.2-158.5°**, **pK<sup>25</sup> 2.90**. Crystallise the acid from water, and dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. The *S*-benzylisothiuronium salt has **m 155-156°** (from EtOH). [Hayes & Brooch *J Am Chem Soc* 65 1577 1943, *Beilstein* 6 H 220, 6 II 211, 6 III 804, 6 IV 1261.]

***p*-Nitrophenyl acetate** [830-03-5] **M 181.2, m 78-79°**. Recrystallise the ester from absolute EtOH [Moss et al. *J Am Chem Soc* 108 5520 1986]. [*Beilstein* 6 IV 1298.]

**2-Nitrophenylacetic acid** [3740-52-1] **M 181.2, m 141-142.5°**, **pK<sup>25</sup> 3.95**. The acid crystallises as yellow needles from EtOH, EtOH/water and dry it over P<sub>2</sub>O<sub>5</sub> under vacuum. The *amide* has **m 160-161°** (from \*C<sub>6</sub>H<sub>6</sub> plates or EtOH, needles). [*Beilstein* 9 III 2282, 9 IV 1687.]

**4-Nitrophenylacetic acid** [104-03-0] **M 181.2, m 153.4-154.6°**, **pK<sup>25</sup> 3.92**. Crystallise the acid from EtOH/water (1:1), then from sodium-dried diethyl ether and dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. The *stypnate* has **m 196.5-197°** (prisms from EtOH). The *amide* has **m 191°** (from EtOH). [*Beilstein* 9 III 2284, 9 IV 1698.]

**4-Nitro-1,2-phenylenediamine** [99-56-9] **M 153.1, m 201°**, **pK<sub>1</sub><sup>25</sup> 1.39 (1-NH<sub>2</sub>)**, **pK<sub>2</sub><sup>25</sup> 2.61 (2-NH<sub>2</sub>)**. Crystallise the diamine from water. [*Beilstein* 13 IV 75.]

**1-(4-Nitrophenyl)ethylamine hydrochloride** [*R*-(+)- 57233-86-0, *S*-( -)- 132873-57-5] **M 202.6, m 225°**, **240-242°(dec)**, **243-245°(dec)**, **248-250°**, **[α]<sub>D</sub><sup>20</sup> (+) and (-) 72° (c 1, 0.05 M NaOH)**, **(+) and (-) 0.3° (H<sub>2</sub>O)**, **pK<sub>Est</sub> ~8.6**. To ensure dryness, the hydrochloride (*ca* 175 g) is extracted with EtOH (3×100ml) and evaporated to dryness (any residual H<sub>2</sub>O increases the solubility in EtOH and lowers the yield). The hydrochloride residue is triturated with absolute EtOH and dried *in vacuo*. The product is further purified by refluxing with absolute EtOH (200 ml for 83g) for 1 hour, and cool to 10° to give 76.6g of hydrochloride **m 243-245°(dec)**. The *free base* is prepared by dissolving in *N* NaOH, extracting with CH<sub>2</sub>Cl<sub>2</sub> (3 x 500ml), drying (Na<sub>2</sub>CO<sub>3</sub>), filtering, evaporating and distilling it. It has **m 27°**, **b 119-120°/0.5mm (105-107°/0.5mm, 157-159°/9mm, d<sub>4</sub><sup>20</sup> 1.1764**, **n<sub>D</sub><sup>20</sup> 1.5688**, **[α]<sub>D</sub><sup>24</sup> ±17.7° (neat)** [Perry et al. *Synthesis* 492 1977, ORD: Nerdel & Liebig *Justus Liebigs Ann Chem* 621 142 1959]. [*Beilstein* 12 IV 2451.]

**4-Nitrophenylhydrazine** [100-16-3] **M 153.1, m 158°(dec)**, **pK<sub>1</sub><sup>25</sup> -9.2 (aqueous H<sub>2</sub>SO<sub>4</sub>)**, **pK<sub>2</sub><sup>25</sup> 3.70**. Crystallise the hydrazine from EtOH. The *hydrochloride* has **m 212°(dec)** (from EtOH, also **m 202-203° dec**).

[*Beilstein* 15 III 331, 15 IV 317.]

**3-Nitrophenyl isocyanate** [3320-87-4] M 164.1, m 52-54°, b 130-131°/11mm. Distil the isocyanate in a vacuum, and/or recrystallise it from petroleum ether (b 28-38°) or toluene/petroleum ether (m 52°). [*Beilstein* 12 H 708, 12 III 1573.]

**4-Nitrophenyl isocyanate** [100-28-7] M 164.1, m 53°, b 137-138°/11mm, 162-164°/20mm. Distil the isocyanate in a vacuum, and/or recrystallise it from petroleum ether (b 28-38°), \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (m 58°) or CCl<sub>4</sub> (m 56-57°). [*Beilstein* 12 H 725, 12 III 1630.]

**2-Nitrophenylpropionic acid** [530-85-8] M 191.1, m 157°(dec), 160.5-161°, 166-167°(dec), pK<sup>25</sup> 2.83 (50% aqueous dioxane), 3.39 (50% aqueous EtOH). Digest the acid with boiling CHCl<sub>3</sub>, then crystallise it from H<sub>2</sub>O. The *amide* has m 159° (plates, from H<sub>2</sub>O). [Schofield & Simpson *J Chem Soc* 516 1945, *Beilstein* 9 H 636, 9 I 267, 9 II 438, 9 III 3067, 9 IV 2330.]

**4-Nitrophenyl trifluoroacetate** [658-78-6] M 235.1, m 37-39°, b 120°/12mm. Recrystallise the ester from CHCl<sub>3</sub>/hexane [Margolis et al. *J Biol Chem* 253 7891 1978]. It sublimes *in vacuo* (m 36-38°). It is *moisture sensitive*. [Sakakibara & Inukai *Bull Chem Soc Jpn* 37 1231 1964.]

**4-Nitrophenyl urea** [556-10-5] M 181.2, m 232°(dec), 238°, 242°. Crystallise the urea from EtOH or hot water. Its UV has λ<sub>max</sub> at 322nm (EtOH). [*Beilstein* 12 II 392, 12 III 1617, 12 IV 1645.]

**3-Nitrophthalic acid** [603-11-2] M 211.1, m 216-218°, pK<sup>25</sup> 3.93. Crystallise 3-nitrophthalic acid from hot water (1.5ml/g). Dry it in air. The *amide* has m 201° (from EtOH). [*Beilstein* 9 H 823, 9 IV 3275.]

**4-Nitrophthalic acid** [610-27-5] M 211.1, m 165°, pK<sup>25</sup> 4.12. Crystallise 4-nitrophthalic acid from Et<sub>2</sub>O, EtOAc or \*C<sub>6</sub>H<sub>6</sub> (m 166°). The *amide* has m 200° (from EtOH). [*Beilstein* 9 H 828, 9 IV 4234.]

**3-Nitrophthalic anhydride** [641-70-3] M 193.1, m 164°. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, Me<sub>2</sub>CO, AcOH, or Ac<sub>2</sub>O (m 164-165°). Dry it at 100°. [*Beilstein* 17 III/IV 6149, 17/11 V 266.]

**4-Nitrophthalic anhydride** [5466-84-2] M 193.1, m 120-121.5°, b 196-197°/8mm. Distil the anhydride in a vacuum and/or recrystallise it from \*C<sub>6</sub>H<sub>6</sub> or Et<sub>2</sub>O/petroleum ether. Dry it *in vacuo*. It forms addition compounds with anthracene (m 118°), and phenanthrene (m 96°). [*Beilstein* 17 III/IV 6150, 17/11 V 267.]

**5-Nitro-2-*n*-propoxyaniline** [553-79-7] M 196.2, m 47.5-48.5°, 49°. pK<sub>Est</sub> ~2.32. Crystallise the aniline from *n*-propyl alcohol/petroleum ether or H<sub>2</sub>O (136mg/L at 20°). [*Beilstein* 13 III 878, 13 IV 897.]

**2-Nitroresorcinol** [601-89-8] M 155.1, m 81-81°, pK<sub>1</sub><sup>20</sup> 6.37, pK<sub>2</sub><sup>20</sup> 9.46. Recrystallise 2-nitroresorcinol from aqueous EtOH. [*Beilstein* 6 H 823.]

**4-Nitrosalicylic acid** [619-19-1] M 183.1, m 235-238°(dec), pK<sup>25</sup> 2.23. Crystallise the acid from H<sub>2</sub>O (m 236-238°) or aqueous EtOH (m 235-236° dec). [*Beilstein* 10 III 194, 10 IV 231.]

**5-Nitrosalicylic acid** [96-97-9] M 183.1, m 233°, pK<sub>1</sub><sup>25</sup> 2.32, pK<sub>2</sub><sup>25</sup> 10.34. Crystallise the acid from Me<sub>2</sub>CO (charcoal), then twice more from Me<sub>2</sub>CO alone, aqueous EtOH (m 234-236°) or H<sub>2</sub>O (m 232-233°). [*Beilstein* 10 III 197, 10 IV 255.]

**Nitrosobenzene** [586-96-9] M 107.1, m 67.5-68°, b 57-59°/18mm. Steam distil nitrosobenzene, then crystallise it from a small volume of EtOH with cooling below 0°, dry it over CaCl<sub>2</sub> in a desiccator at atmospheric pressure, and store it under N<sub>2</sub> at 0°. *Alternatively*, it can be distilled onto a cold finger cooled with brine at ~-10° in a vacuum at 17mm (water pump), while heating in a water bath at 65-70° [Robertson & Vaughan *J Chem Educ* 27 605 1950]. [*Beilstein* 5 IV 702.]

**4-Nitrosodiphenylamine (tautomer of benzoquinone-1,4-phenylimine oxime)** [156-10-5] **M 198.2, m 144-145°(dec), 145.4-146.6°, 144-148°**. The amine forms dark green crystals from EtOH or \*C<sub>6</sub>H<sub>6</sub> (m 143°). It has UV with  $\lambda_{\max}$  at 421nm (EtOH), and it is used for detecting Pd and Rh. *It is highly toxic and a possible carcinogen.* [Beilstein 12 H 207, 12 II 122, 12 III 347, 12 IV 1860ß.]

**1-Nitroso-2-naphthol** [131-91-9] **M 173.2, m 110.4-110.8°, pK<sup>25</sup> 7.63**. Crystallise the naphthol from petroleum ether (b 60-80°, 7.5ml/g). [Beilstein 7 H 712, 7 IV 2419.]

**2-Nitroso-1-naphthol** [132-53-6] **M 173.2, m 162-164°, pK<sup>25</sup> 7.24**. Purify the naphthol by recrystallisation from petroleum ether (b 60-80°) or by dissolving it in hot EtOH, followed by successive addition of small volumes of water (m 158° dec). It also crystallises from \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O. Crystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether gives m (106-109°, also 109.5°). It has  $\lambda_{\max}$  at 274.5 and 382nm (CHCl<sub>3</sub>). It complexes with metals. [Beilstein 7 H 712, 7 I 385, 7 II 647, 7 III 3688, 7 IV 2419.]

**4-Nitroso-1-naphthol** [605-60-7] **M 173.2, m 198°(dec), pK<sup>25</sup> 8.18**. Crystallise the naphthol from \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O (m 197°, dec) and sublime it at 75-80°/0.2mm. It has  $\lambda_{\max}$  at 336nm (EtOH). [Beilstein 7 H 715, 7 II 653, 7 III 3700, 7 IV 2424.]

**2-Nitroso-1-naphthol-4-sulfonic acid (3H<sub>2</sub>O)** [3682-32-4] **M 316.3, m 142-146°(dec), pK<sub>Est</sub> ~6.3 (OH)**. Crystallise the acid from dilute HCl solution. The crystals are dried over CaCl<sub>2</sub> in a vacuum desiccator. It has also been purified by dissolution in aqueous alkali and precipitation by addition of water. It is a reagent for cobalt. [Beilstein 11 H 331, 11 II 189, 11 III 621, 11 IV 668.]

**4-Nitrosophenol (benzoquinone mono oxime)** [104-91-6] **M 123.1, m >124°(dec), pK<sup>25</sup> 6.36**. 4-Nitrosophenol forms yellow crystals from xylene, \*C<sub>6</sub>H<sub>6</sub> (m ~144°) or Et<sub>2</sub>O (m 128-129°, dec). [Beilstein 7 H 622, 7 II 574, 7 III 3367.]

**N-Nitroso-N-phenylbenzylamine** [612-98-6] **M 212.2, m 58°**. Crystallise the amine from absolute EtOH (yellow needles) and dry it in air. It is slightly soluble in organic solvents. [Beilstein 12 H 1071, 12 II 335, 12 III 2335.]

**trans-β-Nitrostyrene** [5153-67-3] **M 149.2, m 60°**. Crystallise the styrene from absolute EtOH, or three times from \*benzene/petroleum ether (b 60-80°) (1:1). [Beilstein 5 III 1180, 5 IV 1352.]

**4-Nitrostyrene** [100-13-0] **M 149.2, m 20.5-21°**. Crystallise it from CHCl<sub>3</sub>/hexane. Purify it by addition of MeOH to precipitate the polymer, then crystallise it at -40° from MeOH. It has also been crystallised from EtOH. [Bernasconi et al. *J Am Chem Soc* 108 4541 1986, Beilstein 5 H 478, 5 III 1180, 5 IV 1351.]

**2-Nitro-4-sulfobenzoic acid** [552-23-8] **M 247.1, m 111°, pK<sub>Est</sub> ~1.65**. Crystallise the acid from dilute HCl. It is *hygroscopic*. [Beilstein 11 H 391, 11 III 685.]

**2-Nitrotoluene** [88-72-2] **M 137.1, m -9.55° (α-form), -3.85° (β-form), b 118°/16mm, 222.3°/760mm, d<sub>4</sub><sup>20</sup> 1.163, n<sub>D</sub><sup>20</sup> 1.545**. Crystallise 2-nitrotoluene (repeatedly) from absolute EtOH by cooling in a Dry-ice/alcohol mixture. Further purify it by passing an alcoholic solution through a column of alumina. [Beilstein 5 IV 845.]

**3-Nitrotoluene** [99-08-1] **M 137.1, m 16°, b 113-114°/15mm, 232.6°, d<sub>4</sub><sup>20</sup> 1.156, n<sub>D</sub><sup>20</sup> 1.544**. Dry 3-nitrotoluene over P<sub>2</sub>O<sub>5</sub> for 24 hours, then fractionally distil it under reduced pressure. [Clark & Taylor *Org. Synth Coll Vol I* 415 1941, Beilstein 5 IV 847.]

**4-Nitrotoluene** [99-99-0] **M 137.1, m 52°**. Crystallise 4-nitrotoluene from EtOH, MeOH/water, EtOH/water (1:1) or MeOH. Dry it in air, then dry it in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. [Wright & Grilliom *J Am Chem Soc* 108 2340 1986, Beilstein 5 IV 848.]

**5-Nitrovanillin (nitroveratric aldehyde)** [6635-20-7] **M 197.2, m 172-175°, 176°, 178°**. It forms yellow

plates from AcOH and needles from EtOH [Slotta & Szyszke *Chem Ber* **68** 184 1935]. With diazomethane, 5-nitro-3,4-dimethoxyacetophenone is formed [Brady & Manjunath *J Chem Soc* **125** 1067 1924]. The *methyl ether* crystallises from EtOAc or AcOH, **m** 88°, 90-91°, and the *phenylhydrazone* has **m** 108-110° (from aqueous EtOH). [Finger & Schott *J Prakt Chem* [2] **115** 288 1927.] The *oxime* has **m** 216° (from EtOH or AcOH), and the *oxime acetate* has **m** 147° (from aqueous EtOH) [Vogel *Monatsh Chem* **20** 384 1899, Brady & Dunn *J Chem Soc* **107** 1861 1915]. [Beilstein **8** III 2064.]

**Nordihydroguaiaretic [1,4-bis(3,4-dihydroxyphenyl)-2,3-dimethylbutane] acid** [500-38-9] **M 302.4, m 184-185°, pK<sub>Est(1)</sub> ~9.7, pK<sub>Est(2)</sub> ~12.** Crystallise the acid from dilute acetic acid. [Beilstein **6** IV 7771.]

**1,2,3,4,6,7,8,9-Octahydroanthracene** [1079-71-6] **M 186.3, m 73°, 73.5°, 78°, d<sub>4</sub><sup>80</sup> 0.9703, n<sub>D</sub><sup>80</sup> 1.5372.** Crystallise the octahydro compound from EtOH, then purify it by zone melting. [Beilstein **5** III 1400.]

***n*-Octylammonium 9-anthranilate** [88020-99-9] **M 351.5, m 134-135°, pK<sup>25</sup> 10.65 (for octylamine).** Recrystallise the ester several times from ethyl acetate.

**4-Octylbenzoic acid** [3575-31-3] **M 234.3, m 99-100°, pK<sup>25</sup> 6.5 (80% aqueous EtOH), pK<sub>Est</sub> ~4.5 (H<sub>2</sub>O).** When crystallised from EtOH it has **m** 139°, but when crystallised from aqueous EtOH it has **m** 99-100°. It forms liquid crystals. [Beilstein **9** H 571, **9** III 2611.]

**4-(*tert*-Octyl)phenol** [140-66-9] **M 206.3, m 85-86°, b 166°/20mm, pK<sub>Est</sub> ~10.4.** Crystallise the phenol from *n*-hexane and/or distil it in a vacuum. [Beilstein **6** III 2051, **6** IV 3484.]

**Opianic acid (2-formyl-4,5-dimethylbenzoic acid)** [519-05-1] **M 210.2, m 150°, pK<sup>25</sup> 3.07.** Crystallise the acid from water. [Beilstein **10** H 990, **10** I 484, **10** II 719, **10** III 4511, **10** IV 3863.]

**Orcine monohydrate (orcinol hydrate, 3,5-dihydroxytoluene)** [6153-39-5] **M 142.2, m 56°, 56-58°, 58°, b 147°/5 mm, pK<sub>1</sub><sup>20</sup> 9.48 (9.26), pK<sub>2</sub><sup>20</sup> 11.20 (11.66).** Purify orcine by recrystallisation from H<sub>2</sub>O as the *monohydrate*. It sublimes *in vacuo*, and the *anhydrous* compound has **m** 106.5-108° (110°, 108°). It can be recrystallised from CHCl<sub>3</sub> (plates) or \*C<sub>6</sub>H<sub>6</sub> (needles or prisms). [UV: Kiss et al. *Bull Soc Chim Fr* 275 1949, Adams et al. *J Am Chem Soc* **62** 732 1940.]

**Orcinol (5-methylresorcinol)** [504-15-4; 6153-39-5 H<sub>2</sub>O] **M 124.2, m 107.5°, m 59-61° (hydrate cf orcine), pK<sub>1</sub><sup>20</sup> 9.36 (9.48), pK<sub>2</sub><sup>20</sup> 11.6 (11.20).** Crystallise orcinol from CHCl<sub>3</sub>/\*benzene (2:3). See *hydrate* in previous entry. [Beilstein **6** H 882, **6** IV 5892.]

**Orthanilic acid (2-aminobenzenesulfonic acid)** [88-21-1] **M 173.2, m >300°(dec), pK<sup>25</sup> 2.49.** Crystallise orthanilic acid from aqueous solution, containing 20ml of conc HCl per L, then crystallise it from distilled water, and dry it in a vacuum desiccator over Sicapent (**m** 315°). When an aqueous solution is chilled below 13.5°, the hydrated form of the acid is obtained. It is used for the determination of nitrite and nitrate. The *S-benzylisothiuronium salt* has **m** 137° (from H<sub>2</sub>O). [Wertheim *Org Synth Coll Vol II* 471 1943, Beilstein **14** H 681, **14** I 714, **14** II 429, **14** III 1896, **14** IV 2638.]

**[2.2]-Paracyclophane (tricyclo[8.2.2.2<sup>4,7</sup>]hexadeca-4,6,10,12,13,15-hexaene)** [1633-22-3] **M 208.3, m 284°, 285-287°, 286-288°, 288-290°.** Purify it by recrystallisation from AcOH. <sup>1</sup>H-NMR δ: 1.62 (Ar-H) and 1.71 (CH<sub>2</sub>) [Waugh & Fessenden *J Am Chem Soc* **79** 846 1957, IR and UV: Cram et al. *J Am Chem Soc* **76** 6132 1954, Cram & Steinberg *J Am Chem Soc* **73** 5691 1951. It complexes with unsaturated compounds: Cram & Bauer *J Am Chem Soc* **81** 5971 1959, Syntheses: Brink *Synthesis* 807 1975, Givens et al. *J Org Chem* **44** 16087 1979, Kaplan et al. *Tetrahedron Lett* 3665 1976]. [Beilstein **5** IV 2223.]

**Para Red (1-(4-nitrophenylazo)-2-naphthol)** [6410-10-2] **M 293.3, m 250-251°, pK<sub>Est</sub> ~9.3.** Crystallise this

dye from AcOH or xylene and dry it *in vacuo*. It has  $\lambda_{\max}$  at 488nm. [*Beilstein* 16 II 70.]

**Pargyline hydrochloride** (Eutonyl, *N*-methyl-*n*-propargylbenzylamine hydrochloride) [306-07-0] **M 195.7, m 154-155°, 155°, pK<sup>25</sup> 6.9**. Recrystallise the salt from EtOH/Et<sub>2</sub>O and dry it *in vacuo*. It is very soluble in H<sub>2</sub>O, in which it is unstable. The *free base* has **b** 101-103°/1mm. It is a glucuronyl transferase inducer and a monoamine oxidase inhibitor. [von Braun et al. *Justus Liebigs Ann Chem* 445 205 1928, Yeh & Mitchell *Experientia* 28 298 1972, Langstrom et al. *Science* 225 1480 1984, *Beilstein* 12 II 548.]

**Pavatrine hydrochloride** [fluorine-9-carboxylic acid, 2-(diethylamino)ethyl ester hydrochloride] [548-65-2] **M 333.7, m 143-144°, pK<sub>Est</sub> ~9.0** Recrystallise the salt from isopropanol, EtOAc/isoPrOH and dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. The *metho-bromide* has **m** 111-117° (from butanone). [*Beilstein* 9 III 3412, 9 IV 2596.]

**Pentabromophenol** [608-71-9] **M 488.7, m 229-230°, pK<sub>Est</sub> ~4.5**. Purify it by crystallisation (charcoal) from toluene then from CCl<sub>4</sub>. Dry it for 2 weeks at *ca* 75°. The *diethylammonium salt* has **m** 191-193° (from MeOH). [*Beilstein* 6 H 206, 6 I 108, 6 II 197, 6 III 766, 6 IV 1069.]

**1-Pentacene** [135-48-8] **M 278.4, m 300°**. It forms blue crystals from \*benzene or nitrobenzene and sublimes in a vacuum. [Clar & John *Ber* 62 940 1929, 64 981 1931, *Beilstein* 5 IV 2721.]

**Pentachloronitrobenzene** [82-68-8] **M 295.3, m 144-145°, 146°**. Crystallise it from EtOH. [*Beilstein* 5 H 247, 5 II 188, 5 III 618.]

**Pentachlorophenol** [87-86-5] **M 266.3, m 190-191°, pK<sup>25</sup> 4.8**. Crystallise it twice from toluene/EtOH. Sublime it *in vacuo*. [*Beilstein* 6 IV 1025.]

**Pentachlorothiophenol** [133-49-3] **M 282.4, m between 228° and 235°, pK<sub>Est</sub> ~1.1**. Crystallise from \*benzene, toluene (**m** 243°) or AcOH (**m** 240° also 242-244°). [*Beilstein* 6 IV 1642.]

**Pentafluorobenzene** [363-72-4] **M 168.1, b 85°/atm, 85-86°/atm, 88-89°/atm, d<sub>4</sub><sup>20</sup> 1.524, n<sub>D</sub><sup>20</sup> 1.3931**. Purify it by distillation and by gas chromatography. Its IR (film) has bands at 1535 and 1512 cm<sup>-1</sup> (\*C<sub>6</sub>H<sub>6</sub> ring). [Stephen & Tatlow *Chem Ind (London)* 821 1957, Nield et al. *J Chem Soc* 166 1959, *Beilstein* 5 IV 639.]

**2,3,4,5,6-Pentafluorobenzoic acid** [602-94-8] **M 212.1, m 101-103°, 104-105°, 106-107°, pK<sup>25</sup> 1.75**. Dissolve the acid in Et<sub>2</sub>O, treat it with charcoal, filter, dry (CaSO<sub>4</sub>), filter again, evaporate and recrystallise the residue from petroleum ether (b 90-100°) after adding a little toluene, to give large colourless plates. Its UV (H<sub>2</sub>O) has  $\lambda_{\max}$  at 265nm ( $\epsilon$  761) (H<sub>2</sub>O). The *S*-benzylisothiuronium salt has **m** 187° (from H<sub>2</sub>O). [McBee & Rapkin *J Am Chem Soc* 73 1366 1951, Nield et al. *J Chem Soc* 166 1959, *Beilstein* 9 IV 956.]

**O-(2,3,4,5,6-Pentafluorobenzyl)hydroxylamine hydrochloride** (PFBHA.HCl) [57981-02-9] **M 249.6, m 215°, 215-216°, pK<sub>Est</sub> ~1.1**. Recrystallise the salt from EtOH to form colourless leaflets. Drying the compound at high vacuum and elevated temperature will result in losses by sublimation. [Youngdale *J Pharm Sci* 65 625 1976, Wehner & Handke *J Chromatog* 177 237 1979, Nambara et al. give incorrect **m** as 115-116° *J Chromatogr* 114 81 1975.]

**2,3,4,5,6-Pentafluorophenol** [771-61-9] **M 184.1, m 33-35°, 38.5-39.5°, b 72-74°/48mm, 142-144°/atm, 143°/atm, n<sub>D</sub><sup>20</sup> 1.4270 (liquid prep), pK<sup>25</sup> 5.53**. It is a *hygroscopic* low melting solid not freely soluble in H<sub>2</sub>O. Purify it by distillation, preferably in a vacuum [Forbes et al. *J Chem Soc* 2019 1959, IR and pKa: Birchall & Haszeldine *J Chem Soc* 13 1959]. IR of a film has  $\nu_{\max}$  3600 (OH) and 1575 (fluoroaromatic breathing) cm<sup>-1</sup>. The *benzoyl* derivative has **m** 74-75°, *3,4-dinitrobenzoyl* derivative has **m** 107°, the *tosylate* has **m** 64-65° (from EtOH) and the *K salt* crystallises from Me<sub>2</sub>CO, **m** 242°(dec), with *1H<sub>2</sub>O-salt* the **m** is 248°(dec) and the *2H<sub>2</sub>O-salt* has **m** 245°(dec). [*Beilstein* 6 IV 782.]

**1-(Pentafluorophenyl)ethanol** [R-(+)- 104371-21-3, S-(-)- 104371-20-2] **M 212.1, m 41-42°, 42°, 42.5-43°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> (+) and (-) 9°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> (+) and (-) 7.5° (c 1, *n*-pentane)**. Recrystallise the ethanol from *n*-pentane at -40°

and sublime it at 25°/0.3mm (use ice-cooled cold finger). It has also been purified by column chromatography through Kieselgel 60 (0.063-0.2mm mesh, Merck) and eluted with EtOAc/*n*-hexane (1:5), then recrystallised from *n*-pentane and sublimed in a vacuum. It has  $R_F$  on Kieselgel 60 F<sub>254</sub> TLC foil and eluting with EtOAc/*n*-hexane (1:5). [Meese *Justus Liebigs Ann Chem* 2004 1986.]

The racemate [75853-08-6] has **m** 32-34°, **b** 77-79°/8mm, 80-82°/37mm,  $n_D^{20}$  1.4426, and the 3,4-dinitrobenzoate has **m** 83° [Nield et al. *J Chem Soc* 166 1959]. [Beilstein 6 IV 3044.]

**Pentamethylbenzene** [700-12-9] **M** 148.3, **m** 53.5-55.1°. Successively crystallise it from absolute EtOH, aqueous EtOH, MeOH, toluene \*C<sub>6</sub>H<sub>6</sub>, and dry it under vacuum. [Rader & Smith *J Am Chem Soc* 84 1443 1962.] It has also been sublimed. The 1,3,5-trinitrobenzene complex (1:1) has **m** 121° (EtOH). Its FT-IR (neat) has  $\nu_{max}$  at 2939.4, 2728.0, 1568.1, 1475.5, 1381.8, 1067.6, 1013.9, 860.7 and 524.6 cm<sup>-1</sup>; [Beilstein 5 H 443, 5 III 1010, 5 IV 1109.]

**Perbenzoic acid** [93-59-4] **M** 138.1, **m** 41-43°, **b** 97-110°/13-15mm,  $pK_{Est}$  ~7.7. Crystallise the peracid from \*benzene or petroleum ether. It sublimes readily and is steam volatile. It is soluble in CHCl<sub>3</sub>, CCl<sub>4</sub> and Et<sub>2</sub>O. [Braun *Org Synth Coll Vol I* 431 1941.] **EXPLOSIVE.**

**Perylene** [198-55-0] **M** 252.3, **m** 273-274°. Purify perylene by silica-gel chromatography of its recrystallised picrate. [Ware *J Am Chem Soc* 83 4374 1961.] Crystallise it from \*benzene, toluene or EtOH and sublime it at 142° in a flow of oxygen-free nitrogen. It forms a 1:1 \*benzene-complex (**m** 223-224.5° needles from \*C<sub>6</sub>H<sub>6</sub>), and a 1:2 \*benzene-complex (**m** 154-155° from \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O). The 2,4,7-trinitrofluoren-9-one has **m** 270-271° (from EtOH/\*C<sub>6</sub>H<sub>6</sub>). [Gorman et al. *J Am Chem Soc* 107 4404 1985, Johansson et al. *J Am Chem Soc* 109 7374 1987, Beilstein 5 III 2521, 5 IV 2689.]

**Phenanthrene** [85-01-8] **M** 178.2, **m** 98°, 98.7-99°, 99.15°, 100.8-101.3°, **b** 148-149°/1mm,  $d_4^{25}$  1.175. Likely contaminants include anthracene, carbazole, fluorene and other polycyclic hydrocarbons. Purify it by distillation from sodium under vacuum, boiling with maleic anhydride in xylene, crystallisation from acetic acid, sublimation and zone melting. It has also been recrystallised repeatedly from EtOH, \*benzene or petroleum ether (**b** 60-70°), with subsequent drying under vacuum over P<sub>2</sub>O<sub>5</sub> in an Abderhalden pistol. Feldman, Pantages and Orchin [*J Am Chem Soc* 73 4341 1951] separated most of the anthracene impurity by refluxing phenanthrene (671g) with maleic anhydride (194g) in xylene (1.25L) under nitrogen for 22 hours, then filtered. The filtrate was extracted with aqueous 10% NaOH, the organic phase was separated, and the solvent was evaporated. The residue, after stirring for 2 hours with 7g of sodium, was distilled in a vacuum, then recrystallised twice from 30% \*benzene in EtOH. It was then dissolved in hot acetic acid (2.2ml/g), and to it was slowly added an aqueous solution of CrO<sub>3</sub> (60g in 72ml H<sub>2</sub>O plus 2.2L of acetic acid), followed by slow addition of conc H<sub>2</sub>SO<sub>4</sub> (30ml). The mixture was refluxed for 15 minutes, diluted with an equal volume of water and cooled. The precipitate was filtered off, washed with water, dried and distilled, then recrystallised twice from EtOH. Further purification is possible by chromatography from a CHCl<sub>3</sub> solution on activated alumina, with \*benzene as eluent, and by zone refining. The picrate (1:1) forms golden yellow needles with **m** 146°, and the styphnate (1:1) has **m** 138-139° (plates or needles from EtOH or EtOH/H<sub>2</sub>O respectively). [Dornfeld et al. *Org Synth Coll Vol III* 134 1955, Beilstein 5 H 667, 5 I 327, 5 II 579, 5 III 2136, 5 IV 2297.]

**Phenanthrene-9-aldehyde** [4707-71-5] **M** 206.3, **m** 102-103°,  $pK^{25}$  -6.39 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise the aldehyde from EtOH and sublime it at 95-98°/0.07mm. The 2,4-dinitrophenylhydrazone has **m** 272-273°. [Beilstein 7 III 2532, 7 IV 1740.]

**9,10-Phenanthrenequinone** [84-11-7] **M** 208.2, **m** 208°,  $pK^{25}$  -7.1 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise the quinone from dioxane or 95% EtOH and dry it under vacuum. [Beilstein 7 IV 2565.]

**Phenethylamine** [64-04-0] **M** 121.2, **b** 87°/13mm,  $d_4^{20}$  0.962,  $n_D^{20}$  1.535,  $pK^{25}$  9.88. Distil the amine from CaH<sub>2</sub>, under reduced pressure, just before use. [Beilstein 12 H 1096, 12 IV 2453.]

**Phenethyl bromide** [103-63-9] **M** 185.1, **b** 92°/11mm,  $d_4^{20}$  1.368,  $n_D^{20}$  1.557. Wash the bromide with conc H<sub>2</sub>SO<sub>4</sub>, water, aqueous 10% Na<sub>2</sub>CO<sub>3</sub> and water again, then dry it with CaCl<sub>2</sub> and fractionally distil it just before

use. [Beilstein 5 IV 907.]

(±)-*N*-1-Phenethyl urea (*N*-α-phenethyl urea) [60295-51-4] M 164.2, m 137°. Crystallise the urea from H<sub>2</sub>O, EtOAc or \*C<sub>6</sub>H<sub>6</sub>. [Buck *J Am Chem Soc* 56 1607 1934, Beilstein 12 I 1096, 12 IV 1440.]

(+)-*S*-*N*-1-Phenethyl urea (*R*-*N*-α-phenethyl urea) [16849-91-5] M 164.2, m 121-122°, [α]<sub>D</sub><sup>25</sup> +48.8° (c 2, EtOH), [α]<sub>D</sub><sup>25</sup> +46.2° (c 4.0, EtOH). Crystallise the (+)-urea from H<sub>2</sub>O or EtOH (m 122-123°). [Marckwald & Meth *Chem Ber* 38 808 1905, Cairns *J Am Chem Soc* 63 871 1941, Beilstein 12 I 1092, 12 III 2398.]

(-)-*S*-*N*-1-Phenethyl urea (*S*-*N*-α-phenethyl urea) [25144-64-3] M 164.2, m 121-122°, [α]<sub>D</sub><sup>25</sup> -43.6° (c 14, EtOH), [α]<sub>D</sub><sup>25</sup> -52.1° (c 3.6, EtOH). Crystallise the (-)-urea from H<sub>2</sub>O or EtOH. [Lovén *J Prakt Chem* [2] 72 313 1905, Beilstein 12 I 1094, 12 III 2398.]

*N*-2-Phenethyl urea [2158-04-5] M 164.2, m 112°. Crystallise the (±)-urea from H<sub>2</sub>O (m 112-113°) or EtOH. The *picrate* has m 113-115° (from H<sub>2</sub>O) [Spica *Gazzetta* 9 567 1879, Shapiro et al. *J Am Chem Soc* 81 2224 1959]. [Beilstein 12 1099, 12 III 2423, 12 IV 2470.]

Phenetole [103-73-1] M 122.2, b 60°/9mm, 77.5°/31mm, 170.0°/760mm, d<sub>4</sub><sup>20</sup> 0.967, n<sub>D</sub><sup>20</sup> 1.50735, n<sub>D</sub><sup>25</sup> 1.50485. Small quantities of phenol can be removed by shaking with NaOH, but this is not a very likely contaminant of commercial material. Fractional distillation from sodium, at low pressures, probably gives adequate purification. It can be dissolved in diethyl ether and washed with 10% NaOH (to remove phenols), then water. The ethereal solution is evaporated, and the phenetole is fractionally distilled under vacuum. [Beilstein 6 H 140, 6 I 80, 6 II 142, 6 III 545.]

Phenocoll hydrochloride (4-ethoxyaniline, *p*-phenetidine HCl) [536-10-6] M 230.7, m 234°, pK<sup>28</sup> 5.20. Crystallise the salt from water then sublime it *in vacuo*. [Beilstein 13 IV 1017.]

Phenol [108-95-2] M 94.1, m 40.9°, b 85.5-86.0°/20mm, 180.8°/760mm, d<sub>4</sub><sup>20</sup> 1.06, n<sub>D</sub><sup>41</sup> 1.54178, n<sub>D</sub><sup>46</sup> 1.53957, pK<sup>25</sup> 9.86 (10.02). Steam is passed through a boiling solution containing 1mole of phenol and 1.5-2.0moles of NaOH in 5L of H<sub>2</sub>O until all non-acidic material has distilled. The residue is cooled, acidified with 20% (v/v) H<sub>2</sub>SO<sub>4</sub>, and the phenol is separated, dried with CaSO<sub>4</sub> and fractionally distilled under reduced pressure. It is then fractionally crystallised several times from its melt [Andon et al. *J Chem Soc* 5246 1960]. Purification *via* the benzoate has been used by Berliner, Berliner and Nelidow [*J Am Chem Soc* 76 507 1954]. The *benzoate*, (m 70°, b 314°/760mm), is crystallised from 95% EtOH, then hydrolysed to the free phenol by refluxing with two equivalents of KOH in aqueous EtOH until the solution becomes homogeneous. It is acidified with HCl and extracted with diethyl ether. The ether layer is freed from benzoic acid by thorough extraction with aqueous NaHCO<sub>3</sub>, and, after drying and removing the ether, the phenol is distilled.

Phenol has also been crystallised from a 75% w/w solution in water by cooling to 11° and seeding with a crystal of the hydrate. The crystals are centrifuged off, rinsed with cold water (0-2°), saturated with phenol, and dried. It can be crystallised from petroleum ether [Berasconi & Paschalis *J Am Chem Soc* 108 2969 1986].

Draper and Pollard [*Science* 109 448 1949] added 12% water, 0.1% aluminium (can also use zinc) and 0.05% NaHCO<sub>3</sub> to phenol, and distilled it at atmospheric pressure until the azeotrope was removed. The phenol was then distilled at 25mm. Phenol has also been dried by distillation from the \*benzene solution to remove the water/\*benzene azeotrope and the excess \*benzene, followed by distillation of the phenol at reduced pressure under nitrogen. Processes such as this are probably adequate for analytical grade phenol which has as its main impurity water. Phenol has also been crystallised from petroleum ether/\*benzene or petroleum ether (b 40-60°). The purified material is stored in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or CaSO<sub>4</sub>. [Beilstein 6 IV 531.]

Phenol-2,4-disulfonic acid [96-77-5] M 254.2, pK<sub>1</sub> <1, pK<sub>2</sub> <1, pK<sub>3</sub> ~8.3. Crystallise the acid from EtOH/diethyl ether. [Beilstein 11 H 250, 11 I 58, 11 II 139, 6 III 522.]

Phenolphthalein [77-09-8] M 319.2, m 263°, pK<sub>Est(1)</sub> ~ 4.2, pK<sub>Est(2)</sub> ~ 9.8. Dissolve it in EtOH (7ml/g), then dilute it with eight volumes of cold water, filter and heat on a water-bath to remove most of the alcohol and the phenolphthalein that precipitates is filtered off and dried *in vacuo*. [Beilstein 18 II 119, 18 III/IV 1945, 18/4 V

188.]

**Phenolphthalol** [81-92-5] **M 306.3, m 201-202°, pK<sub>Est</sub> ~ 9.8.** Crystallise it from aqueous EtOH. [Beilstein 6 H 1146, 6 II 1110, 6 IV 7623.]

**Phenoxyacetic acid** [122-59-8] **M 152.2, m 98-99°, pK<sup>25</sup> 3.18.** Crystallise the acid from water or aqueous EtOH. [Beilstein 6 IV 634.]

**Phenoxyacetyl chloride** [701-99-5] **M 170.6, b 112°/10mm, 102°/16mm, 225-226°/atm, d<sub>4</sub><sup>20</sup> 1.235, n<sub>D</sub><sup>20</sup> 1.534.** If it has no OH band in the IR then distil it in a vacuum, taking precautions for the moisture-sensitive compound. If it contains free acid (due to hydrolysis, OH bands in the IR), then add an equal volume of redistilled SOCl<sub>2</sub>, reflux for 2-3 hours, evaporate and distil the residue in a vacuum as before. The *amide* has **m 101°.** [McElvain & Carney *J Am Chem Soc* 68 2592 1946, Beilstein 6 III 613.]

**4-Phenoxyaniline** [139-59-3] **M 185.2, m 95°, pK<sup>20</sup> 4.44 (50% aqueous EtOH).** Crystallise 4-phenoxyaniline from water. [Beilstein 13 IV 1020.]

**Phenoxybenzamine** [*N*-(2-chloroethyl)-*N*-(1-methyl-2-phenoxyethyl)benzylamine] [59-96-1] **M 303.5, m 38-40°, hydrochloride** [63-92-3] **M 340.0, m 137.5-140°, pK<sub>Est</sub> ~4.2.** The free base is crystallised from petroleum ether, and the *HCl* is crystallised from EtOH/diethyl ether. [Beilstein 12 IV 2204.]

**2-Phenoxybenzoic acid** [2243-42-7] **M 214.2, m 113°, b 355°/760mm, pK<sup>15/25</sup> 3.53.** Crystallise the acid from aqueous EtOH or H<sub>2</sub>O (**m 114°**). [Beilstein 10 H 65, 10 I 28, 10 II 40, 10 III 99, 10 IV 132.]

**3-Phenoxybenzoic acid** [3739-38-6] **M 214.2, m 145°, pK<sup>25</sup> 3.95.** Crystallise the acid from aqueous EtOH. [Beilstein 10 H 138, 10 III 247, 10 IV 316.]

**Phenoxybutyric acid** [6303-58-8] **M 180.2, m 64°, 65-66°, 82-83°, 99°, b 180-185°/12mm, pK 3.17.** It has been purified by recrystallisation from petroleum ether, \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O/petroleum ether, EtOH and from H<sub>2</sub>O. It can be steam distilled or distilled in a good vacuum. [UV: Ramart-Lucas & Hoch *Bull Soc Chim Fr* [4] 51 824 1932, Dann & Arndt *Justus Liebigs Ann Chem* 587 38 1954.] The *acid chloride* has **b 154-156°/20mm** [Hamford & Adams *J Am Chem Soc* 57 921 1935], and the *amide* crystallises from \*C<sub>6</sub>H<sub>6</sub> as needles with **m 113°.** [Beilstein 6 IV 645.]

**2-Phenoxypropionic acid (lactic acid *O*-phenylether)** [940-31-8] **M 166.2, m 115-116°, b 105-106°/5mm, 265-266°/758mm, pK<sup>25</sup> 3.11.** Crystallise the acid from water. [Beilstein 6 H 163, 6 II 158, 6 III 614.]

**Phensuximide (*N*-methyl-2-phenylsuccinimide)** [86-34-0] **M 189.2, m 71-73°.** Crystallise phensuximide from hot 95% EtOH (**m 72-73°**). At 25° 1g of the imide dissolves in 1g of \*C<sub>6</sub>H<sub>6</sub>, 18g of Et<sub>2</sub>O, 9.5g of EtOH, 5.1g of MeOH and 235g of H<sub>2</sub>O. [Beilstein 21 II 300, 21 III/IV 5465.]

**Phenylacetamide** [103-81-1] **M 135.2, m 158.5°.** Crystallise the acetamide repeatedly from absolute EtOH, EtOAc (**m 160-161°**) or H<sub>2</sub>O (**m 159-160°**). Dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub>. [Beilstein 9 H 347, 9 III 2193, 9 IV 1632.]

**Phenyl acetate** [122-79-2] **M 136.2, b 78°/10mm, d<sub>4</sub><sup>20</sup> 1.079, n<sub>D</sub><sup>22</sup> 1.5039.** Phenyl acetate acid is freed from phenol and acetic acid by washing (either directly or as a solution in pentane) with aqueous 5% Na<sub>2</sub>CO<sub>3</sub>, then with saturated aqueous CaCl<sub>2</sub>, drying with CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, and fractionally distilling under reduced pressure. [Beilstein 6 II 153, 6 III 595, 6 IV 611.]

**Phenylacetic acid** [103-82-2] **M 136.2, m 76-77°, b 140-150°/20mm, pK<sub>1</sub> -7.59 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub> 4.31.** Crystallise the acid from petroleum ether (b 40-60°), isopropyl alcohol, 50% aqueous EtOH or hot water (**m 77.8-78.2°**). Dry it *in vacuo*. It can be distilled under a vacuum. [Beilstein 9 II 294, 9 III 2169.]

**Phenylacetone (1-phenylpropan-2-one)** [103-79-9] **M 134.2, b 69-71°/3mm,  $d_4^{20}$  1.00,  $n_D^{20}$  1.516.** Convert the ketone to the *semicarbazone* and crystallise it three times from EtOH (**m** 186-187°). The *semicarbazone* is then hydrolysed with 10% phosphoric acid, and the ketone is distilled. [Kumler et al. *J Am Chem Soc* **72** 1463 1950, *Beilstein* **7** H 303, **7** I 161, **7** II 223, **7** III 1037, **7** IV 687.]

**4'-Phenylacetophenone** [92-91-1] **M 196.3, m 120.3-121.2°, b 196-210°/18mm, 325-327°/760mm.** Crystallise it from EtOH or acetone. It can also be distilled under reduced or atmospheric pressure. The *semicarbazone* has **m** 131-132° (aqueous EtOH). [*Beilstein* **7** H 443, **7** III 2134, **7** IV 1407.]

**Phenylacetylene** [536-74-3] **M 102.1, b 75°/80mm,  $d_4^{20}$  0.930,  $n_D^{25}$  1.5463, pK ~19.** Distil phenylacetylene through a spinning band column. It should be filtered through a short column of alumina before use [Collman et al. *J Am Chem Soc* **108** 2988 1986; for pK see Brandsma *Preparative Acetylenic Chemistry*, 1st Edn Elsevier 1971, p. 15, ISBN 0444409475]. [*Beilstein* **5** IV 1525.]

**Phenylalaninol (2-amino-3-phenylpropan-1-ol)** [*R*-(+)- 5267-64-1, *S*-(-)- 3182-95-4] **M 151.2, m 91-92°, 91.5°, 92-94°, b 80°/11mm (Kügelrohr),  $[\alpha]_{546}^{20}$  (+) and (-) 28°,  $[\alpha]_D^{20-25}$  (+) and (-) 23-28.7° (c 1-5, EtOH), pK<sub>Est</sub> ~9.3.** It can be recrystallised from Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 40-60°) or toluene and distilled in a vacuum. It has been purified by dissolving in Et<sub>2</sub>O, drying over K<sub>2</sub>CO<sub>3</sub>, filtering, evaporating to a small volume, cooling in ice and collecting the plates. Store them in the presence of KOH (i.e. CO<sub>2</sub>-free atm). [Karrer & Ehrhardt *Helv Chim Acta* **34** 3203 1951, Oeda *Bull Chem Soc Jpn* **13** 465 1938.] The *picrate* has **m** 141-141.5° (from EtOH/petroleum ether). The *hydrogen oxalate* has **m** 177°, 161-162° [Hunt & McHale *J Chem Soc* 2073 1957]. The *racemate* has **m** 87-88° from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (75-77° from Et<sub>2</sub>O), and the *hydrochloride* has **m** 139-141° [Fodor et al. *J Chem Soc* 1858 1951]. [*Beilstein* **13** IV 1920.]

**3-Phenylallyl chloride (cinnamyl chloride)** [*E*: 18685-01-3, *Z*: 39199-93-4] **M 152.6, trans: m 7-8°, b 92-93°/3mm,  $d_4^{25}$  1.086,  $n_D^{25}$  1.5802, cis: 85°/3mm,  $d_4^{25}$  1.0891,  $n_D^{25}$  1.5746.** Distil the chloride under vacuum three times from K<sub>2</sub>CO<sub>3</sub>. [Hatch & Alexander *J Am Chem Soc* **72** 643 1950, *Beilstein* **5** III 1186.]

**Phenyl 4-aminosalicylate (Phenamisal)** [133-11-9] **M 229.2, m 153°, pK<sub>Est(1)</sub> ~2.0 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~9.7 (OH).** Crystallise the ester from EtOH (**m** 155°, also 149-150.5°), aqueous EtOH (**m** 147-149°), or isopropanol. It is **tuberculostatic**. [*Beilstein* **13** IV 1979.]

**4-Phenylanisole (4-methoxybiphenyl)** [361-37-6] **M 184.2, m 89.9-90.1°.** Crystallise the biphenyl from \*benzene/petroleum ether. Dry it under vacuum in an Abderhalden pistol. It has  $\lambda_{\max}$  at 259.5nm (hexane). [*Beilstein* **6** H 674, **6** II 625, **6** III 3321, **6** IV 4600.]

**9-Phenylanthracene** [602-55-1] **M 254.3, m 153-154°, 156°, b 417°/760mm.** Chromatograph it on alumina in \*C<sub>6</sub>H<sub>6</sub> and recrystallise it from AcOH or toluene. [*Beilstein* **5** H 725, **5** II 639, **5** III 2462.]

***p*-Phenylazobenzoyl chloride** [104-24-5] **M 244.7, m 93°.** Crystallise the acid chloride from petroleum ether (b 60-80°). [*Beilstein* **16** III 224.]

**4-Phenylazo-1-naphthylamine** [131-22-6] **M 247.3, m 125-125.5°.** Crystallise the dye from cyclohexane or aqueous EtOH. [Brode et al. *J Am Chem Soc* **74** 4641 1952, *Beilstein* **16** H 361, **16** III 406, **16** IV 546.]

**4-Phenylazophenacyl bromide** [62625-24-5] **M 317.3, m 103-104°.** Purify the bromide on a column silica gel, using petroleum ether/Et<sub>2</sub>O (9:1 v/v) as solvent. It forms orange yellow crystals from \*C<sub>6</sub>H<sub>6</sub> (**m** 113-114°) or petroleum ether/Et<sub>2</sub>O (**m** 114.5-115°). [*Beilstein* **16** II 22, **16** III 53, **16** IV 79.]

**4-Phenylazophenol (4-hydroxyazobenzene)** [1689-82-3] **M 198.2, m 155°, pK<sub>1</sub><sup>25</sup> -0.93, pK<sub>2</sub><sup>25</sup> 8.2.** Crystallise the dye from \*benzene or 95% EtOH. [*Beilstein* **16** II 38, **16** III 86, **16** IV 159.]

**Phenyl benzenethiosulfonate (diphenyldisulfoxide)** [1212-08-4] **M 250.3, m 36-37°, 45-46°, 45-47°.** Recrystallise the disulfoxide from EtOH or MeOH. It has also been purified from phenylsulfide impurities by

dissolving in  $\text{CHCl}_3$ , washing with aqueous saturated  $\text{NaHCO}_3$ , drying ( $\text{Na}_2\text{SO}_4$ ), filtering, evaporating the filtrate, and the residual oil is passed through a silica gel column (600g) and eluted with hexane/\* $\text{C}_6\text{H}_6$  (1L, 4:1, eluting PhSSPh) then \* $\text{C}_6\text{H}_6$  (1L) which elutes PhSSO<sub>2</sub>Ph. [Trost & Massiot *J Am Chem Soc* **99** 4405 1977, Knoevenagel & Römer *Chem Ber* **56** 215 1923, *Beilstein* **11** IV 220.]

**Phenyl benzoate** [93-99-2] **M 198.2, m 69.5°**, **b 198-199°**. Crystallise the ester from EtOH using *ca* twice the volume needed for complete dissolution at 69°. [*Beilstein* **9** IV 303.]

**Phenyl-1,4-benzoquinone** [363-03-1] **M 184.2, m 114-115°**. Crystallise the quinone from heptane, petroleum ether (b 60-70°), \* $\text{C}_6\text{H}_6$  (m 113.5-114.5°) or EtOH (m 112-113°) and sublime it *in vacuo*. [Carlson & Miller *J Am Chem Soc* **107** 479 1985, *Beilstein* **7** H 740, **7** III 3764.]

**1-Phenylbiguanide** [102-02-3] **M 177.2, m 144-146°**, **pK<sub>1</sub><sup>32</sup> 2.16, pK<sub>2</sub><sup>32</sup> 10.74**. Crystallise the biguanide from water or toluene. [*Beilstein* **12** H 370, **12** I 236, **12** III 807.]

**N-Phenyl-bis(trifluoromethanesulfonimide)** [ditriflic phenylimide, *N,N*-bis(trifluoromethylsulfonyl) aniline] [37595-74-7] **M 357.3, m 93-94°**, **100-102°**. This ditriflic imide is a very stable triflating agent. If hydrolysis appears to have occurred; it would be better to prepare anew from purified aniline with 2 mols of redistilled trifluoroacetic anhydride in  $\text{CH}_2\text{Cl}_2$  at -78°, with 2 mols of  $\text{Et}_3\text{N}$  in order to keep the aniline as the free base and to mop up the free acid formed. It is then evaporated *in vacuo* to give the crystalline reagent. [Murry & Scott *Tetrahedron Lett* **24** 979 1983, Crisp & Scott *Synthesis* 335 1985, Hendrickson & Bergeron *Tetrahedron Lett* 4607 1973, Mascarenas et al. *Tetrahedron* **47** 3485 1991.]

**S(-)-1-Phenylbutanol** [22135-49-5] **M 150.2, m 46-47°**, **46-48°**, **49°**, **b 90-92°/2mm**. [ $\alpha$ ]<sub>D</sub><sup>18</sup> -51.4° (c 5,  $\text{CHCl}_3$ ), -44.7° (c 5.13, \* $\text{C}_6\text{H}_6$ ). Purify the alcohol by distillation, and the distillate crystallises on cooling. The hydrochloride has [ $\alpha$ ]<sub>D</sub><sup>20</sup> +45.1° (c 4.8, \* $\text{C}_6\text{H}_6$ ). The (-)-hydroperoxide has b 58°/0.005mm, n<sub>D</sub><sup>20</sup> 1.5123,  $\alpha$ <sub>D</sub><sup>18</sup> -2.14°, (l = 0.5dcm, neat). [Holding & Ross *J Chem Soc* 145 1954, Davies & Feld *J Chem Soc* 4637 1958.] The (±)-racemate has b 73°/0.05mm, and its 4-nitrophenylhydrazone has m 58°. [*Beilstein* **6** IV 3272.]

**2-Phenylbutyramide** [90-26-6] **M 163.2, m 86°**, **87°**. Crystallise the amide from  $\text{H}_2\text{O}$ , EtOH,  $\text{Et}_2\text{O}$ /petroleum ether or \* $\text{C}_6\text{H}_6$ . [*Beilstein* **9** I 212, **9** II 356, **9** III 2466.]

**2-Phenylbutyric acid** [*R*(-)- 938-79-4, *S*(+)- 4286-15-1] **M 164.2, b 102-104°/760mm**, **d<sub>4</sub><sup>20</sup> 1.056**, **n<sub>D</sub><sup>20</sup> 1.521**, [ $\alpha$ ]<sub>D</sub><sup>20</sup> (+) and (-) 96° (c 2.5, \* $\text{C}_6\text{H}_6$ ), [ $\alpha$ ]<sub>D</sub><sup>23</sup> (-) and (+) 5.8° (neat), **pK<sub>Est</sub> ~4.3**. Purify the acids by distillation at atmospheric pressure using an efficient column. The acid chlorides have b 106-107°/20mm, [ $\alpha$ ]<sub>D</sub><sup>18</sup> (-) and (+) 108° (c 2, \* $\text{C}_6\text{H}_6$ ). [Levene et al. *J Biol Chem* **100** 589 1933, Gold & Aubert *Helv Chim Acta* **41** 1512 1958, ORD in heptane: Rothen & Levene *J Chem Phys* **7** 975 1939, *Beilstein* **9** III 2461.]

**3-Phenylbutyric acid** [*R*(-)- 772-14-5, *S*(+)- 772-15-6] **M 164.2, b 94-95°/3mm**, **134°/4mm**, **d<sub>4</sub><sup>26</sup> 1.066**, **n<sub>D</sub><sup>25</sup> 1.5167**, [ $\alpha$ ]<sub>D</sub><sup>20</sup> (-) and (+) 57° (c 1, \* $\text{C}_6\text{H}_6$ ), **pK<sub>25</sub> 4.40**. Purify the acids as the 2-isomer above, i.e. by distillation, but under a good vacuum. [Prelog & Scherrer *Helv Chim Acta* **42** 2227 1959, Levene & Marker *J Biol Chem* **93** 761 1932, **100** 685 1933, Cram *J Am Chem Soc* **74** 2137 1952.] The *R*-amide crystallises from  $\text{H}_2\text{O}$ , with m 101.5-102°, and [ $\alpha$ ]<sub>D</sub><sup>20</sup> -16.5° (c 1.2, EtOH). The racemic acid has m 39-40°, b 134-136°/6mm, 158°/12mm [Marvel et al. *J Am Chem Soc* **62** 3499 1940]. [*Beilstein* **9** IV 1813.]

**4-Phenylbutyric acid** [1821-12-1] **M 164.2, m 50°**, **pK<sub>25</sub> 4.76**. Crystallise the acid from petroleum ether (b 40-60°). [*Beilstein* **9** IV 1811.]

**O-Phenyl chlorothionoformate** [1005-56-7] **M 172.6, b 81-83°/6mm**, **91°/10mm**, **d<sub>4</sub><sup>20</sup> 1.276**, **n<sub>D</sub><sup>20</sup> 1.585**. Purify it by dissolving in  $\text{CHCl}_3$ , washing with  $\text{H}_2\text{O}$ , drying ( $\text{CaCl}_2$ ), filtering, evaporating and distilling twice under a vacuum to give a clear yellow liquid. **It is reactive and POISONOUS—work in a fume cupboard.** Store it in sealed ampoules under  $\text{N}_2$ . A possible impurity is *O,O'*-diphenyl thiocarbonate which has m 106° and remains behind in the distilling flask. [Bögemann et al. in *Methoden Der Organischen Chemie (Houben-Weyl)* 4th edn (E. Müller Ed.) **Vol 9 Schwefel-Selen-Tellur Verbindungen** pp. 807-808 1955, Rivier & Schalch *Helv*

*Chim Acta* **6** 612 1923, *Kalson Chem Ber* **20**, 2384 1987, *Rivier & Richard Helv Chim Acta* **8** 490 1925, *Schönberg & Varga Justus Liebigs Ann Chem* **483** 176 1930, *Schönberg & Vargha Chem Ber* **64** 1390 1931, *Beilstein* **6** III 609.]

**Phenyl cinnamate** [2757-04-2] **M 224.3, m 75-76°, b 205-207°/15mm.** Crystallise the cinnamate from EtOH (2ml/g). It can also be distilled under reduced pressure. [Womack & McWhirter *Org Synth Coll Vol III* 715 1955, *Beilstein* **9** H 583, **9** II 387, **9** III 2689, **9** IV 2011.]

**$\alpha$ -Phenylcinnamic acid (2,3-diphenylprop-2-enoic acid)** [*cis-E* 91-48-5, *trans-Z* 91-47-4] **M 224.3, m 174°(cis), m 138-139°(trans), pK<sup>25</sup> 4.8 (60% aqueous EtOH).** Crystallise the acid from Et<sub>2</sub>O/petroleum ether. Crystallise the *cis-isomer* from petroleum ether or EtOH (**m 174°**) and has a pK<sup>25</sup> of 4.44, and the *cis-amide* from aqueous Me<sub>2</sub>CO (**m 174°**). Crystallise the *trans-isomer* from Et<sub>2</sub>O/petroleum ether or EtOH (**m 140°**), and the *trans-amide* from CHCl<sub>3</sub>/petroleum ether (**m 167-168°**). [*Beilstein* **6** H 691, **9** III 3414.]

***o*-Phenylenediamine** [95-54-5] **M 108.1, m 100-101°, pK<sub>1</sub><sup>25</sup> 0.67 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>25</sup> 4.47 (4.85).** Crystallise the diamine from aqueous 1% sodium hydrosulfite (charcoal), wash it with ice-water and dry it in a vacuum desiccator, or sublime it *in vacuo*. It has been purified by recrystallisation from toluene and zone refined [Anson et al. *J Am Chem Soc* **108** 6593 1986]. Purification by refluxing a CH<sub>2</sub>Cl<sub>2</sub> solution containing charcoal is also carried out followed by evaporation and recrystallisation [Koola & Kochi *J Org Chem* **52** 4545 1987], protect from light. The *acetate* has **m 186°**. [*Beilstein* **13** IV 38.]

***m*-Phenylenediamine** [108-45-2] **M 108.1, m 61-63°, 62-63°, 62.5°, 63-64°, b 146°/22mm, 282-284°/760mm, 284-287°/atm, d<sub>10</sub><sup>10</sup> 1.1422, n<sub>D</sub><sup>57.7</sup> 1.6340, pK<sub>1</sub><sup>25</sup> 2.41, pK<sub>2</sub><sup>25</sup> 4.98.** Purify the diamine by distillation under a vacuum followed by recrystallisation from EtOH (rhombs) and if necessary redistillation. It should be protected from light; otherwise it darkens rapidly. [Neilson et al. *J Chem Soc* 371 1962, IR: Katritzky & Jones *J Chem Soc* 3674, 2058 1959, UV: Forbes & Leckie *Can J Chem* **36** 1371 1958.] The *hydrochloride* has **m 277-278°**, and the *bis-4-chlorobenzenesulfonyl* derivative has **m 220-221°** from H<sub>2</sub>O (214-215°, from MeOH/H<sub>2</sub>O) [Runge & Pfeiffer *Chem Ber* **90** 1737 1957]. The *acetate* has **m 191°**. [*Beilstein* **13** IV 79.]

***p*-Phenylenediamine** [106-50-3] **M 108.1, m 140°, pK<sub>1</sub><sup>25</sup> 2.89, pK<sub>2</sub><sup>25</sup> 6.16.** Crystallise the diamine from EtOH or \*benzene, and sublime it *in vacuo*; protect it from light. The *acetate* has **m 304°**. [*Beilstein* **13** IV 104.]

***o*-Phenylenediamine dihydrochloride** [615-28-1] **M 181.1, m 180°.** Crystallise the salt from dilute HCl (60ml conc HCl, 40ml water, with 2g stannous chloride), after treatment of the hot solution with charcoal by adding an equal volume of conc HCl and cooling in an ice-salt mixture. The crystals are washed with a small amount of conc HCl and dried in a vacuum desiccator over NaOH. [*Beilstein* **13** IV 38.]

**1,4-Phenylene diisothiocyanate (bitoscanate)** [4044-65-9] **M 192.3, m 129-131°, 130-131°, 132°.** Purify bitoscanate by recrystallisation from AcOH, petroleum ether (b 40-60°), Me<sub>2</sub>CO or aqueous Me<sub>2</sub>CO. [van der Kerk et al. *Rec Trav Chim Pays Bas* **74** 1262 1955, Leiber & Slutkin *J Org Chem* **27** 2214 1962, *Beilstein* **13** IV 174.]

**1-Phenyl-1,2-ethanediol** [*R*-( -) 16355-00-3, *S*-(+) 25779-13-9] **M 138.2, m 64-67°, 65-66°, [α]<sub>D</sub><sup>24</sup> (-) and (+) 40.5° (c 2.8, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> (-) and (+) 39° (c 3, EtOH).** Purify the diol by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/ligroin and sublime it at 1-2mm. [Arpesella et al. *Gazetta* **85** 1354 1955, Prelog et al. *Helv Chim Acta* **37** 221 1954, *Beilstein* **6** IV 5939.]

***dl*-1-Phenylethanol (*dl*- $\alpha$ -methylbenzyl alcohol)** [98-85-1; 13323-81-4] **M 122.2, b 60.5-61.0°/3mm, 106-107°/22-23mm, d<sub>4</sub><sup>20</sup> 1.01, n<sub>D</sub><sup>25</sup> 1.5254.** Purify the alcohol *via* its hydrogen phthalate. [See Houssa & Kenyon *J Chem Soc* 2260 1930.] Shake it with a solution of ferrous sulfate, and the alcohol layer is washed with distilled H<sub>2</sub>O, dried (MgSO<sub>4</sub>) and fractionally distilled. [*Beilstein* **6** II 444.]

**2-Phenylethanol** [60-12-8] **M 122.2, b 215-217°, d 1.020.** Purify the ethanol by shaking it with a solution of ferrous sulfate, and the alcohol layer is washed with distilled water and fractionally distilled. [*Beilstein* **6** IV

3067.]

**Phenyl ether (diphenyl ether)** [101-84-8] **M 170.2, m 27.0°, b 83-84°/1mm, 138°/21mm, 257°/760mm, d<sub>4</sub><sup>20</sup> 1.074, n<sub>D</sub><sup>30.7</sup> 1.57596.** Crystallise the ether from 90% EtOH. Melt it, wash it with 3M NaOH and water, dry it with CaCl<sub>2</sub> and fractionally distil it under reduced pressure. Fractionally recrystallise it from its melt and store over P<sub>2</sub>O<sub>5</sub>. [Beilstein 6 IV 562.]

**1-Phenylethyl isocyanate (α-methylphenyl isocyanate)** [*R*-(+)- 33375-06-3, *S*(-)- 14649-03-7] **M 147.2, b 82-83°/12-14mm, d<sub>4</sub><sup>20</sup> 1.045, n<sub>D</sub><sup>20</sup> 1.513, [α]<sub>D</sub><sup>24</sup> (+) and (-) 2° (c 3.5, \*C<sub>6</sub>H<sub>6</sub>), (+) and (-) 10.5° (neat).** Purify the isocyanates by fractional distillation under a vacuum. With ammonia they give the *ureido* derivatives which crystallise from H<sub>2</sub>O with **m 121-122°, [α]<sub>D</sub><sup>25</sup> (+) and (-) 48.8°.** [Cairns *J Am Chem Soc* 63 870 1941.] The *racemate* has **b 90-94°/3mm, 96°/18mm** [Seifitan *Justus Liebigs Ann Chem* 562 75 1949]. [Beilstein 12 IV 2443.]

**(±)-*p*-α-Phenylethylphenol** [1988-89-2] **M 198.3, m 56.0-56.3°, 64°, b 126°/0.4mm, 165-170°/5mm, 315-316°/742.2mm, pK<sub>Est</sub> ~10.3.** Crystallise the phenol from petroleum ether. *S*-(+)-*enantiomer* has [α]<sub>D</sub> +10.3° (\*C<sub>6</sub>H<sub>6</sub>). [Okamoto et al. *Bull Chem Soc Jpn* 39 303 1966.]

**5-(α-Phenylethyl)semioxamazide** [93-95-8] **M 207.1, m 167-168° (*l*-), 157° (*dl*-).** Crystallise it from EtOH. The *l*-*enantiomer* has [α]<sub>D</sub><sup>25</sup> -102.5° (c 1, CHCl<sub>3</sub>). [Leonard & Boyer *J Org Chem* 15 42 1950.]

**9-Phenyl-3-fluorone (2,6,7-trihydroxy-9-phenylxanthen-3-one)** [975-17-7] **M 320.3, m >300°(dec), 350°, λ<sub>max</sub> 462nm (ε 4.06 x 10<sup>4</sup>, in 1M HCl aqueous EtOH).** Recrystallise it from warm, acidified EtOH by addition of ammonia. The crude material (1g) can be extracted with EtOH (50ml) in a Soxhlet apparatus for 10 hours to remove impurities. Impurities can be detected by paper electrophoresis. The *triacetate* forms yellow needles from EtOH (**m 230-233°**). [Petrova et al. *Anal Lett* 5 695 1972, *Beilstein* 18 H 199, 18 I 404, 18 III/IV 2824.]

**Phenylhydrazine** [100-63-0] **M 108.1, m 23°, b 71.8°/1mm, 137-138°/18mm, 241-242°/760mm, d<sub>4</sub><sup>20</sup> 1.10, n<sub>D</sub><sup>20</sup> 1.607, pK<sub>1</sub><sup>20</sup> -5.2 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>25</sup> 5.27.** Purify phenylhydrazine by chromatography, then crystallise it from petroleum ether (b 60-80°)\*benzene. Store it in the dark under N<sub>2</sub> as it turns yellow, then red, on exposure to air. It is best stored as the *hydrochloride salt*; see below. [Coleman *Org Synth Coll Vol I* 442 1941, Shaw & Stratton *J Chem Soc* 5004 1962, *Beilstein* 15 IV 50.]

**Phenylhydrazine hydrochloride** [59-88-1] **M 144.5, m 244°, 250-254°(dec).** Dissolve 100g of phenylhydrazine hydrochloride in 200ml of warm H<sub>2</sub>O (60-70°) during 1-3 hours, then add 1L of boiling EtOH. The solution is filtered, while still hot, through Whatman No 2 filter paper and cooled in a refrigerator. The precipitate is collected on a medium sintered-glass filter and recrystallised twice this way, then washed with cold EtOH, dried thoroughly and stored in a stoppered brown bottle. [Peterson et al. *Anal Chem* 29 144 1957.] Hough, Powell and Woods [*J Chem Soc* 4799 1956] boiled the hydrochloride with three times its weight of water, filtered hot (charcoal), added one-third volume of concentrated HCl and cooled to 0°. The crystals were washed with acetone, and dried over P<sub>2</sub>O<sub>5</sub> under vacuum. The salt has also been crystallised from 95% EtOH, and it can be sublimed. [Coleman *Org Synth Coll Vol I* 442 1941, *Beilstein* 15 III 71.]

**Phenylhydroxylamine (*N*-hydroxyaniline)** [100-65-2] **M 109.1, m 81°, 82°, 83-84°, pK<sup>25</sup> 3.2.** Impure base deteriorates rapidly. Crystallise it from H<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (40-60°). The *picrate* has **m 186°** (from EtOH), and the *benzenesulfonate salt* has **m 70°** (dec) (EtOH/\*C<sub>6</sub>H<sub>6</sub>). [Beilstein 15 H 2, 15 I 3, 15, II 4, 15 III 5, 15 IV 4.]

**2-Phenyl-1,3-indandione** [83-12-5] **M 222.2, m 149-151°, pK<sup>20</sup> 4.12 (1% aqueous MeOH).** Crystallise the dione from EtOH (**m 156°**) or CHCl<sub>3</sub> (**m 148-150°**). [Beilstein 7 H 808, 7 III 4100, 7 IV 2570.]

**Phenylisocyanate** [103-71-9] **M 119.1, b 45-47°/10mm, d<sub>4</sub><sup>20</sup> 1.093, n<sub>D</sub><sup>20</sup> 1.536.** Distil phenylisocyanate under reduced pressure from P<sub>2</sub>O<sub>5</sub>. [Beilstein 12 IV 864.]

**Phenylisothiocyanate (phenyl mustard oil)** [103-72-0] **M 135.2, m -21°, b 95°/12mm, 117.1°/33mm, 221°/760mm,  $d_4^{25}$  1.1288,  $n_D^{23.4}$  1.64918.** It is insoluble in H<sub>2</sub>O, but soluble in Et<sub>2</sub>O and EtOH. If impure (due to formation of thiourea), then steam distil it into a receiver containing 5-10ml of N H<sub>2</sub>SO<sub>4</sub>. Separate the oil, dry over CaCl<sub>2</sub> and distil it under vacuum. [Dains et al. *Org Synth Coll Vol I* 447 1941, *Beilstein* 12 IV 867.]

**8-Phenylmenthol** [1R,2S,5R-(-)- 65253-04-5, 1S,2R,5S-(+)- 57707-91-2] **M 232.4,  $[\alpha]_D^{20}$  (-) and (+) 26° (c 2, EtOH).** Dissolve the menthol in toluene, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and chromatograph it on a silica gel column and eluting with 5% Et<sub>2</sub>O in petroleum ether to give an oil with the desired rotation. Its IR has  $\nu_{\max}$  at 3420cm<sup>-1</sup> (OH) with consistent <sup>1</sup>H NMR [Corey & Ensley *J Org Chem* 43 1610 1978, Whitesell et al. *Tetrahedron* 42 2993 1986, Bednarski & Danishefsky *J Am Chem Soc* 108 7060 1986].

**Phenyl methanesulfonate** [16156-59-5] **M 172.1, m 61-62°.** Crystallise the sulfonate from MeOH or from H<sub>2</sub>O (m 61.5°). [*Beilstein* 6 H 176, 6 III 650, 6 IV 689.]

**Phenyl methanesulfonyl fluoride (PMSF)** [329-98-6] **M 174.2, m 90-91°, 92-93°.** Purify PMSF by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or CHCl<sub>3</sub>/petroleum ether. [Davies & Dick *J Chem Soc* 483 1932, cf Tullock & Coffman *J Org Chem* 23 2016 1960.] It is a general protease inhibitor (specific for trypsin and chymotrypsin) and is a good substitute for diisopropylphosphorofluoridate [Fahrney & Gould *J Am Chem Soc* 85 997 1963]. [*Beilstein* 11 III 331.]

**2-Phenylnaphthalene** [612-94-2] **M 204.3, m 102-103°, 103.5°, 103-104°, b 185-190°/5mm, 357-358°/760mm.** Chromatograph it on alumina in \*benzene and crystallise it from aqueous EtOH or MeOH/EtOH. It has been sublimed. It has  $\lambda_{\max}$  at 248 and 286nm (methylcyclohexane). The 2,4,7-trinitrofluoren-9-one has m 169.5-170.5° (from EtOH/\*C<sub>6</sub>H<sub>6</sub>). [*Beilstein* 5 H 687, 5 II 603, 5 III 2231, 5 IV 2412.]

**N-Phenyl-1-naphthylamine** [90-30-2] **M 219.3, m 63.7-64.0°, pK<sub>Est</sub> ~0.1.** Crystallise it from EtOH, petroleum ether or \*C<sub>6</sub>H<sub>6</sub>/EtOH. Dry it under vacuum in an Abderhalden pistol. [*Beilstein* 12 H 1224.]

**N-Phenyl-2-naphthylamine** [135-88-6] **M 219.3, m 107.5-108.5°, 110°, pK<sub>Est</sub> ~0.5.** Crystallise it from EtOH, MeOH, glacial acetic acid or \*benzene/hexane. [*Beilstein* 12 H 1275, 12 I 535, 12 II 716, 12 III 2991.]

**4-Phenylphenacyl bromide** [135-73-9] **M 275.2, m 126°.** Crystallise (charcoal) the bromide from EtOH (15ml/g), or ethyl acetate/petroleum ether (b 90-100°). [*Beilstein* 7 III 2137.]

**2-Phenylpropanal** [93-53-8] **M 134.2, b 206°/760mm,  $d_4^{20}$  1.001,  $n_D^{20}$  1.5183.** It may contain up to 15% of acetophenone. Purify it *via* the bisulfite addition compound [Lodge & Heathcock *J Am Chem Soc* 109 3353 1987] and see Chapter 2 for preparation and decomposition of bisulfite adducts. [*Beilstein* 7 IV 695.]

**Phenylpropionic acid** [637-44-5] **M 146.2, m 137.8-138.4°, pK<sup>25</sup> 2.23.** Crystallise the acid from \*benzene, CCl<sub>4</sub> (m 136°) or aqueous EtOH. The *S*-benzylisothiuronium salt has m 184-186° (from EtOH). [*Beilstein* 9 II 436, 9 III 3061, 9 IV 2327.]

**RS-2-Phenylpropionic acid** [492-37-5] **M 150.2, m 16-16.5°, b 153-155°/20mm, 189°/48mm, 260-262°/760mm,  $d_4^{20}$  1.10,  $n_D^{20}$  1.522, pK<sup>25</sup> 4.3.** Fractionally distil the acid, or recrystallise it from petroleum ether (b 40-60°) with strong cooling (see references below). [*Beilstein* 9 II 348.]

**2-Phenylpropionic acid** [R-(-)- 7782-26-5, S-(+)- 7782-24-3] **M 150.2, m 30.3-31°, 30-32°, b 115°/1-2mm, 142°/12mm,  $[\alpha]_D^{20}$  (-) and (+) 99.7° (l = 1 dcm, neat), (-) and (+) 89.1° (c 1.7, EtOH), (-) and (+) 75° (c 1.6, CHCl<sub>3</sub>).** Purify the acids by vacuum distillation and by recrystallisation from petroleum ether. The *S*-anilide has m 103-104° (from H<sub>2</sub>O or CHCl<sub>3</sub>/\*C<sub>6</sub>H<sub>6</sub>),  $[\alpha]_D^{25}$  +47° (c 9, Me<sub>3</sub>CO) [Argus & Kenyon *J Chem Soc* 916 1939, Campbell & Kenyon *J Chem Soc* 25 1946, Levene et al. *J Biol Chem* 88 27, 34 1930]. [*Beilstein* 9 III 2417, 9 IV 1779.]

**3-Phenylpropionic acid (hydrocinnamic acid)** [501-52-0] **M 150.2, m 48-48.5°, pK<sup>25</sup> 4.56.** Crystallise the acid from \*benzene, CHCl<sub>3</sub> or petroleum ether (b 40-60°). Dry it in a vacuum. [Beilstein 9 H 508.]

**3-Phenylpropyl bromide** [637-59-2] **M 199.1, b 110°/12mm, 128-129°/29mm, d<sub>4</sub><sup>20</sup> 1.31.** Wash the bromide successively with conc H<sub>2</sub>SO<sub>4</sub>, water, 10% aqueous Na<sub>2</sub>CO<sub>3</sub> and again with water, then dry it with CaCl<sub>2</sub> and fractionally distil it just before use. [Beilstein 5 IV 982.]

**Phenylpyruvic acid** [156-06-9] **M 164.2, m 150-154°, 158-159°, pK<sub>Est</sub> ~2.1.** Recrystallise the acid from \*C<sub>6</sub>H<sub>6</sub>. The *phenylhydrazone* has **m 173°** [Zeller *Helv Chim Acta* 26 1614 1943, Hopkins & Chisholm *Can J Research* [B] 24 89 1946]. The *2,4-dinitrophenylhydrazone* has **m 162-164° (189°, 192-194°)** [Fones *J Org Chem* 17 1952]. [Beilstein 10 IV 2760.]

**Phenyl salicylate (Salol)** [118-55-8] **M 214.2, m 41.8-42.6°, pK<sub>Est</sub> ~9.9.** Fractionally crystallise salol from its melt, then crystallise it from \*benzene. [Beilstein 10 IV 154.]

**3-Phenylsalicylic acid** [304-06-3] **M 214.3, m 186-187.5°, pK<sub>Est(1)</sub> ~2.8 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~11.0 (OH).** Dissolve the acid in *ca* 1 equivalent of saturated aqueous Na<sub>2</sub>CO<sub>3</sub>, filter and precipitate it by adding 0.8 equivalents of M HCl. Crystallise it from ethylene dichloride (charcoal), and sublime it at 0.1mm. [Brooks et al. *J Chem Soc* 661 1961.]

**1-Phenylsemicarbazide** [103-03-7] **M 151.2, m 172°, 173.5°.** Crystallise it from water and dry it in a vacuum over KOH. [Beilstein 15 H 287, 15 II 106, 15 III 184, 15 IV 180.]

**4-Phenylsemicarbazide** [537-47-3] **M 151.2, m 122°.** Crystallise it from water and dry it in a vacuum over KOH. [Beilstein 12 II 221, 12 III 822.]

**Phenylsuccinic acid** [*R*-(--)- 46292-93-7, *S*-(+)- 4036-30-1] **M 194.2, m 173-176°, 178.5-179°, 179-180°, [α]<sub>D</sub><sup>25</sup> (-) and (+) 171° (c 2, Me<sub>2</sub>CO), [α]<sub>D</sub><sup>26-30</sup> (-) and (+) 148° (c 0.27-5, EtOH), pK<sub>1</sub><sup>25</sup> 3.78, pK<sub>2</sub><sup>25</sup> 5.55.** Purify the acids by re-precipitation from alkali and recrystallisation from H<sub>2</sub>O. [Naps & Johns *J Am Chem Soc* 62 2450 1940, Fredga & Matell *Bull Soc Chim Belg* 62 47 1953, Wren & Williams *J Chem Soc* 109 572 1916.] The *racemate* [635-51-8] has **m 166-168°, 168°** after recrystallisation from H<sub>2</sub>O or MeCN. Its *S*-benzylisothiuronium salt has **m 164-165° (from EtOH)** [Griediger & Pedersen *Acta Chem Scand* 9 1425 1955]. [Beilstein 9 IV 3351.]

**1-Phenylthiosemicarbazide** [645-48-7] **M 167.2, m 200-201°(dec).** Crystallise it from EtOH. [Beilstein 15 IV 183.]

**4-Phenylthiosemicarbazide** [5351-69-9] **M 167.2, m 140°.** Crystallise it from EtOH. [Beilstein 12 IV 827.]

**1-Phenyl-2-thiourea** [103-85-5] **M 152.1, m 154°.** Crystallise the thiourea from water and dry it at 100° in air. [Beilstein 12 IV 804.]

**Phenyltoloxamine [2-(2-dimethylaminoethoxy)-diphenylmethane] hydrochloride** [6152-43-8] **M 291.8, m 119-120°, pK<sup>25</sup> 9.3 (free base).** Crystallise the salt from isobutyl methyl ketone. The *free base* has **b 144°/1mm.** [Beilstein 6 III 3351, 6 IV 4630.]

**Phenyl 4-toluenesulfonate** [640-60-8] **M 248.2, m 94.5-95.5°, 95.8-97.8°.** Crystallise the ester from MeOH or glacial acetic acid. [Beilstein 11 H 99, 11 II 47, 11 III 200, 6 IV 271.]

**Phenyl 4-tolylcarbonate** [13183-20-5] **M 228.2, m 94°.** Purify the carbonate by preparative GLC with 20% Apiezon on Embacel, and sublime it *in vacuo*. [Beilstein 6 H 398.]

**1-Phenyl-2,2,2-trifluoroethanol** [*R*-( -)- 10531-50-7, *S*-(+)- 340-06-7] **M 176.1, b 74-76°/10mm, 125-127°/760mm, d<sub>4</sub><sup>20</sup> 1.301, n<sub>D</sub><sup>20</sup> 1.4632, [α]<sub>D</sub><sup>20</sup> (-) and (+) 31° (neat).** Purify the chiral alcohols by fractional

distillation preferably in a vacuum: [Morrison & Ridgeway *Tetrahedron Lett* 573 1969, NMR: Pirkle & Beare *J Am Chem Soc* 90 6250 1968.] The racemate [340-05-6] has **b** 52-54°/2mm, 57-59°/2mm, 64-65°/5mm,  $d_4^{20}$  1.293,  $n_D^{20}$  1.457, and the 2-carbobenzoyl derivative has **m** 137-138° [Mosher et al. *J Am Chem Soc* 78 4374 1956]. [Beilstein 6 IV 3043.]

**Phenylurea** [64-10-8] **M** 136.2, **m** 148°, **pK**<sup>25</sup> -1.45 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise the urea from boiling water (10ml/g) or amyl alcohol (**m** 149°). Dry it in a steam oven at 100°. The 1:1 *resorcinol complex* has **m** 115° (from EtOAc/\*C<sub>6</sub>H<sub>6</sub>). [Beilstein 12 H 346, 12 II 204, 12 III 760, 12 IV 734.]

**Phloroacetophenone (2H<sub>2</sub>O) (2',4',6'-trihydroxyacetophenone)** [480-66-0] **M** 186.2, **m** 218-219°, **pK**<sub>Est(1)</sub>~7.9, **pK**<sub>Est(2)</sub>~12.0. Crystallise the ketone from hot H<sub>2</sub>O (35ml/g). [Beilstein 8 IV 2729.]

**Phloretin [2',4',6'-trihydroxy-3-(*p*-hydroxyphenyl)propiophenone]** [60-82-2] **M** 274.3, **m** 264-271°(dec), **pK**<sub>Est(1)</sub>~7.5, **pK**<sub>Est(2)</sub>~8.0, **pK**<sub>Est(3)</sub>~10, **pK**<sub>Est(4)</sub>~12 (phenolic OH's). Crystallise phloretin from aqueous EtOH. [Zemplén & Bognár *Chem Ber* 75 1040 1942, Zemplén *Chem Ber* 76 386 1943, Beilstein 8 IV 3518.]

**Phloroglucinol (2H<sub>2</sub>O) (benzene-1,3,5-triol)** [6099-90-7 (2H<sub>2</sub>O), 108-73-6 (anhydrous)] **M** 126.1, **m** 217-219°, 117° (anhydrous), **pK**<sub>1</sub><sup>25</sup> -7.74 (HClO<sub>4</sub>), **pK**<sub>2</sub><sup>20</sup> 7.97, **pK**<sub>3</sub><sup>20</sup> 9.23. Crystallise the triol from water, and store it in the dark under nitrogen. [Beilstein 6 IV 7361.]

***o*-Phthalic acid** [88-99-3] **M** 166.1, **m** 211-211.5°, **pK**<sub>1</sub><sup>25</sup> 2.76 (3.05), **pK**<sub>2</sub><sup>25</sup> 4.92 (4.73). Crystallise phthalic acid from water. [Beilstein 9 IV 3167.]

**Phthalic anhydride** [85-44-9] **M** 148.1, **m** 132°, **b** 295°. Distil the anhydride under reduced pressure. Purify it from the acid by extracting with hot CHCl<sub>3</sub>, filtering and evaporating. The residue is crystallised from CHCl<sub>3</sub>, CCl<sub>4</sub> or \*benzene, or sublimed. Fractionally crystallise it from its melt. Dry it under vacuum at 100°. [Saltiel *J Am Chem Soc* 108 2674 1986, Beilstein 17/11 V 253.]

**Phthalide** [87-41-2] **M** 134.1, **m** 72-73°, **pK** -7.98 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise phthalide from water (75ml/g) and dry it in air on filter paper. [Beilstein 17/10 V 7.]

**Phthalimide** [85-41-6] **M** 147.1, **m** 235°, 238°, **pK** 8.30. Crystallise the imide from EtOH (20ml/g) (charcoal), or sublime it. For potassium phthalimide see entry in "Metal-organic Compounds", Chapter 5. [Beilstein 21/10 V 270.]

**Phthalonitrile (1,2-dicyanobenzene)** [91-15-6] **M** 128.1, **m** 141°. Crystallise the nitrile from EtOH, toluene or \*benzene. It has also been distilled under high vacuum. It is steam volatile. [Beilstein 9 H 815, 9 II 602, 9 III 4199, 9 IV 3268.]

**Phthalylsulfacetamide** [131-69-1] **M** 362.3, **m** 304°. It is prepared by hydrolysis of *N*-acetamidophenylphthalimide by boiling for 3 hours in 15% aqueous KOH followed by treatment with Norit, filtration and acidification with AcOH. Crystallise the phthalamide from hot H<sub>2</sub>O or EtOH [Jain et al. *J Indian Chem Soc* 24 174 1947]. It is **antibacterial**. [Beilstein 14 III 2073.]

**Phthiocol (2-hydroxy-3-methylnaphtha-1,4-quinone)** [483-55-6] **M** 188.1, **m** 173-174°, **pK**<sub>Est</sub> ~4.2. Crystallise the quinone from diethyl ether/petroleum ether. [Beilstein 8 III 2568, 8 IV 2375.]

**Physodic acid [4,4',6'-trihydroxy-6-(2-oxoheptyl)-2'-pentyl-2,3'-oxydibenzoic acid 1,5-lactone]** [84-24-2] **M** 470.5, **m** 205°, **pK**<sub>Est(1)</sub>~3.0, **pK**<sub>Est(2)</sub>~10, **pK**<sub>Est(3)</sub>~13. Crystallise the acid from MeOH. The *diacetate* has **m** 155-156° (from Me<sub>2</sub>CO/CS<sub>2</sub>). [Beilstein 19 II 329, 19 III/IV 3988.]

**Picene** [213-14-3] **M** 278.3, **m** 364°, 366-367°, 367-369°. Crystallise picene from isopropylbenzene/xylene. After sublimation at 300°/2mm followed by crystallisation from xylene (charcoal), it gives white glistening leaflets with **m** 366-366.5° [Newman *J Org Chem* 9 524 1944]. The 2,4,7-trinitrofluoren-9-one has **m** 257-

257.8° (from \*C<sub>6</sub>H<sub>6</sub>). [Beilstein 5 H 735, 5 I 369, 5 III 2555, 5 IV 2724.]

**Picric acid** [88-89-1] M 229.1, m 122-123°, pK<sup>25</sup> 0.33 (0.37). Crystallise the acid first from acetic acid, then acetone, toluene, CHCl<sub>3</sub>, aqueous 30% EtOH, 95% EtOH, MeOH or H<sub>2</sub>O. Dry it in a vacuum for 2 hours. Alternatively, dry it over Mg(ClO<sub>4</sub>)<sub>2</sub> or fuse (CARE) and allow it to solidify under a vacuum three times. Because it is **EXPLOSIVE**, picric acid should be stored moistened with H<sub>2</sub>O, and only small portions should be dried at any one time. The dry acid should **NOT** be heated. [Beilstein 6 IV 1388.]

**Picryl chloride (2-chloro-1,3,7-trinitrobenzene)** [88-88-0] M 226.3, m 83°. Crystallise the chloride from CHCl<sub>3</sub> or EtOH. The 2:1 \*C<sub>6</sub>H<sub>6</sub>-complex has m 39° (from \*C<sub>6</sub>H<sub>6</sub>). [Beilstein 5 II 205, 5 III 645, 5 IV 757.]

**Picryl iodide (2-iodo-1,3,7-trinitrobenzene)** [4436-27-5] M 340.0, m 164-165°. Crystallise the iodide from \*benzene. [Beilstein 5 III 647, 5 IV 758.]

**Pinacyanol chloride (Quinaldine Blue)** [2768-90-3] M 388.9, m 270°(dec). Crystallise the chloride from EtOH/diethyl ether. [Beilstein 23 I 90.]

**Piperic acid [trans,trans-5-(3,4-methylenedioxyphenyl)-2,4-pentadieneoic acid]** [136-72-1] M 218.2, m 217°, pK<sub>Est</sub> ~4.7. Crystallise the acid from EtOH. It turns yellow in light. It sublimes with partial decomposition. It has UV with λ<sub>max</sub> at 340nm (MeOH). [Beilstein 19 H 281, 19 II 30, 19 III/IV 3565.]

**Piperonal** [120-57-0] M 150.1, m 37°, b 140°/15mm, 263°/760mm. Crystallise piperonal from aqueous 70% EtOH or EtOH/water. [Beilstein 19/4 V 225.]

**Piperonylic acid** [94-53-1] M 166.1, m 229°, pK<sup>25</sup> 4.50. Crystallise the acid from EtOH or water. [Beilstein 19/7 V 300.]

**Polystyrene (PS)** [9003-53-6]. Precipitate polystyrene repeatedly from CHCl<sub>3</sub> or toluene solution by addition of MeOH. Dry it *in vacuo*. [Miyasaka et al. *J Phys Chem* 92 249 1988.]

**Procaine** [4-(2-diethylaminomethoxycarbonyl)aniline] [59-46-1] M 236.3, m 51° (dihydrate), 61° (anhydrous), pK<sub>1</sub><sup>15</sup> 2.45, pK<sub>2</sub><sup>15</sup> 8.91. Procaine crystallises as the *dihydrate* from aqueous EtOH and as the *anhydrous* material from petroleum ether or diethyl ether. The latter is *hygroscopic*. [Beilstein 14 IV 1138.]

**p-(1-Propenyl)phenol** [*cis/trans* 6380-21-8, 539-12-8] M 134.2, m 93-94°, pK<sub>Est</sub> ~10.2. Crystallise the phenol from water. [Beilstein 6 III 2394, 6 IV 3796.]

**1-Propyl-3-(p-chlorobenzenesulfonyl) urea** [94-20-2] M 260.7, m 126-128°, 127-129°. Crystallise the urea from aqueous EtOH. [Beilstein 11 IV 119.]

**n-Propyl gallate** [121-79-9] M 212.2, m 150°. Crystallise the ester from aqueous EtOH or \*C<sub>6</sub>H<sub>6</sub> (m 146-146.5°). [Beilstein 10 III 2078, 10 IV 2003.]

**Protocatechualdehyde (3,4-dihydroxybenzaldehyde)** [139-85-5] M 138.1, m 153°. Crystallise the aldehyde from water or toluene and dry it in a vacuum desiccator over KOH pellets or shredded wax respectively. [Beilstein 8 IV 1762.]

**1S,2S-Pseudoephedrine (1-hydroxy-1-phenyl-2-methylaminopropane)** [90-82-4] M 165.2, m 118-119°, [α]<sub>D</sub><sup>20</sup> +53.0° (EtOH), +40.0° (H<sub>2</sub>O), pK<sup>25</sup> 9.71. Crystallise the amine from dry diethyl ether, or from water and dry it in a vacuum desiccator. [Beilstein 13 IV 1878.]

**1S,2S-Pseudoephedrine hydrochloride** [345-78-8] M 210.7, m 181-182°, 185-188°, [α]<sub>D</sub><sup>20</sup> +61° (c 1 H<sub>2</sub>O). Crystallise the salt from EtOH. [Beilstein 13 IV 1878.]

**Purpurin (1,2,4-trihydroxy-5,10-anthraquinone)** [81-54-9] **M 256.2, m 253-256°**,  $\text{pK}_{\text{Est}(1)} \sim 7.0$  (2-OH),  $\text{pK}_{\text{Est}(2)} \sim 9.0$  (4-OH),  $\text{pK}_{\text{Est}(3)} \sim 11.1$  (1-OH). Crystallise purpurin from aqueous EtOH, dry it at 100°. [Beilstein 8 IV 3568.]

**Pyrene (benzo[def]phenanthrene)** [129-00-0] **M 202.3, m 149-150°**. Crystallise pyrene from EtOH, glacial acetic acid, \*benzene or toluene. Purify it also by chromatography of CCl<sub>4</sub> solutions on alumina, with \*benzene or *n*-hexane as eluent. [Backer & Whitten *J Phys Chem* 91 865 1987.] It can also be zone refined and purified by sublimation. Marvel and Anderson [*J Am Chem Soc* 76 5434 1954] refluxed pyrene (35g) in toluene (400ml) with maleic anhydride (5g) for 4 days, then added 150ml of aqueous 5% KOH and refluxed for 5 hours with occasional shaking. The toluene layer was separated, washed thoroughly with H<sub>2</sub>O, concentrated to about 100ml and allowed to cool. Crystalline pyrene was filtered off and recrystallised three times from EtOH or acetonitrile. [Chu & Thomas *J Am Chem Soc* 108 6270 1986, Russell et al. *Anal Chem* 50 2961 1986.] The material is free from anthracene derivatives. Another purification step involves passage of pyrene in cyclohexane through a column of silica gel. It can be sublimed in a vacuum and zone refined. The *picrate* has **m 224°**. [Kano et al. *J Phys Chem* 89 3748 1985, Beilstein 5 IV 2467.]

**Pyrene-1-aldehyde** [3029-19-4] **M 230.3, m 125-126°**. Recrystallise the aldehyde three times from aqueous EtOH. [Beilstein 7 IV 1821.]

**1-Pyrenebutyric acid** [3443-45-6] **M 288.4, m 184-186°**,  $\text{pK}_{\text{Est}} \sim 4.1$ . Crystallise the butyric acid from \*benzene, EtOH, EtOH/water (7:3 v/v) or \*C<sub>6</sub>H<sub>6</sub>/AcOH. Dry it over P<sub>2</sub>O<sub>5</sub>. [Chu & Thomas *J Am Chem Soc* 108 6270 1986, Beilstein 9 IV 2731.]

**1-Pyrenecarboxylic acid** [19694-02-1] **M 230.3, m 274-274°**,  $\text{pK}_{\text{Est}} \sim 3.2$ . Crystallise the acid from \*C<sub>6</sub>H<sub>6</sub>, chlorobenzene, nitrobenzene or 95% EtOH. [Beilstein 9 H 712, 9 III 3575.]

**1-Pyrenesulfonic acid** [26651-23-0] **M 202.2, m >350°**,  $\text{pK}_{\text{Est}} < 0$ . Crystallise the sulfonic acid from EtOH/H<sub>2</sub>O. The *sulfonyl chloride* has **m 120°(dec)**. [Vollmann et al. *Justus Liebigs Ann Chem* 531 32 1937 and *Justus Liebigs Ann Chem* 540 189 1939, Beilstein 11 H 198, 11 III 448.]

**1,3,6,8-Pyrenetetrasulfonic acid** [6528-53-6] **M 522.2, m >400°**,  $\text{pK}_{\text{Est}} < 0$ . Crystallise the tetrasulfonic acid from water. The tetra-Na salt crystallises also from H<sub>2</sub>O. [Tietz & Bayer *Justus Liebigs Ann Chem* 540 189 1939, Beilstein 11 III 486.]

***p*-Quaterphenyl** [135-70-6] **M 306.4, m 312-314°**, **b 428°/18mm**. Recrystallise *p*-quaterphenyl from dimethyl sulfoxide at *ca* 50°. [Beilstein 5 II 669.]

**Quinalizarin (1,2,5,8-tetrahydroxy-9,10-anthraquinone)** [81-61-8] **M 272.2, m 275°**,  $\text{pK}_{\text{Est}(1)} \sim 7.1$  (1-OH),  $\text{pK}_{\text{Est}(2)} \sim 9.9$  (8-OH),  $\text{pK}_{\text{Est}(3)} \sim 11.1$  (5-OH),  $\text{pK}_{\text{Est}(4)} \sim 11.8$  (2-OH). Crystallise the quinone from acetic acid or nitrobenzene. It can be sublimed *in vacuo*. The Cu salt forms red crystals. [Beilstein 8 H 549, 8 I 755, 8 II 584.]

**Quinhydrone (1:1 complex of hydroquinone and benzoquinone)** [106-34-3] **M 218.2, m 168°, 167-172°**. Crystallise quinhydrone from H<sub>2</sub>O at 65°, then dry it *in vacuo*. [Beilstein 7 H 617, 7 IV 2069.]

**Quinizarin (1,4-dihydroxy-9,10-anthraquinone)** [81-64-1] **M 240.2, m 200-202°**,  $\text{pK}_1^{25} 9.90$  (9.5),  $\text{pK}_2^{25} 11.18$ . Crystallise quinizarin from glacial acetic acid. [Beilstein 8 H 450, 8 IV 3260.]

***p*-Quinquephenyl (*p*-pentaphenyl)** [61537-20-0] **M 382.5, m 388.5°**. Recrystallise *p*-pentaphenyl from boiling dimethyl sulfoxide (**b** 189°, lowered to 110°). The solid obtained on cooling is filtered off and washed repeatedly with toluene, then with conc HCl. The final material is washed repeatedly with hot EtOH. It is also recrystallised from pyridine, then sublimed *in vacuo*. [Campbell & McDonald *Org Synth Coll Vol V* 986 1973, Beilstein 5 II 709, 5 III 2680, 5 IV 2874.]

**Resorcinol** [108-46-3] **M 110.1, m 111.2-111.6°, pK<sub>1</sub><sup>25</sup> 9.23, pK<sub>2</sub><sup>25</sup> 13.05.** Crystallise resorcinol from \*benzene, toluene or \*benzene/diethyl ether. The *benzoate* has **m 117°**. [Beilstein 6 IV 2069.]

**Retene (1-methyl-7-isopropylphenanthrene)** [483-65-8] **M 234.3, m 89°, 90°, 99°, 100.5-101°, b 208°/10mm.** Crystallise retene from EtOH. The *picrate* has **m 126°** (from EtOH). The 1:1 *1,3,5-trinitrobenzene-complex* has **m 144°** (from EtOH). [Beilstein 5 H 683, 5 II 598, 5 III 2199.]

**ParaRosaniline (4,4',4''-triaminotriptyllium [triphenylmethane] carbonium ion, para-fuschin, paramagenta)** [467-62-9] **M 305.4, pK<sup>25</sup> 7.57 and free base has pK >13.** Dissolve the dye in EtOH (1.16g in 30ml), filter and add aqueous NH<sub>3</sub> till neutral (colourless) and precipitate it by adding H<sub>2</sub>O giving 0.8g, **m 247°(dec, sintering at 230°)**. Dissolve it in EtOH, neutralise with NH<sub>3</sub> till colourless, add 0.1g of charcoal, filter, and repeat. then add H<sub>2</sub>O (100ml) to precipitate the colourless *carbinol* (pseudo-base) and dry it *in vacuo*, **m 257°(dec, also 205°, sintering at 232°)**. [Weissberger & Theile *J Chem Soc* 148 1934.] The *carbinol* is slightly soluble in H<sub>2</sub>O but is soluble in acids (e.g. HCl to give the coloured *chloride* [569-61-9]) and EtOH [pK: Goldacre & Phillips *J Chem Soc* 172 1949]. The *perchlorate* (dark red with a green shine) has **m 300°** and explodes at 317° [Dilthey & Diaklage *J Prakt Chem* [2] 129 1931]. [Beilstein 13 IV 2283.]

**Rosaniline HCl (Magenta I, Fuschin)** [632-99-5] **M 337.9, m >200°(dec).** Purify the dye by dissolving it in EtOH, filtering and adding H<sub>2</sub>O. Filter or centrifuge it and wash the precipitate with Et<sub>2</sub>O and dry it in air. It has also been recrystallised from water and dried *in vacuo* at 40°. The crystals have a metallic green lustre. It has UV with  $\lambda_{\max}$  in EtOH at 543nm ( $\epsilon$  93,000). Its solubility in H<sub>2</sub>O is 0.26%. A carmine red colour is obtained in EtOH. It is paraRosaniline with a methyl group. [Scalan *J Am Chem Soc* 57 887 1937.]

***p*-Rosolic acid (4-[bis-{4-hydroxyphenyl}methylene]-2,5-cyclohexadien-one, 4',4''-di-hydroxy-fuschon, aurin, corallin)** [603-45-2] **M 290.3, m 292°, 295-300° (dec with liberation of phenol), 308-310°(dec), pK<sub>1</sub> 3.11, pK<sub>2</sub> 8.62.** It forms green crystals with a metallic luster, but the colour depends on the solvent used. When recrystallised from brine (saturated aqueous NaCl) acidified with HCl, it forms red needles, but when recrystallised from EtO/AcOH, the crystals have a beetle iridescent green colour. It has been recrystallised from Me<sub>2</sub>CO (although it dissolves slowly), methyl ethyl ketone, 80-95% AcOH and from AcOH/\*C<sub>6</sub>H<sub>6</sub>. An aqueous KOH solution is golden yellow, and a 70% H<sub>2</sub>SO<sub>4</sub> solution is deep red in colour. An *alternative* purification is to dissolve this triphenylmethane dye in 1.5% of aqueous NH<sub>3</sub>, filter, and heat to 70-80°, then acidify with dilute AcOH by adding it slowly with vigorous stirring, whereby the aurin separates as a brick-red powder or as purplish crystals depending on the temperature and period of heating. Filter off the solid, wash it with H<sub>2</sub>O and a little dilute AcOH, then H<sub>2</sub>O again. Stir this solid with Et<sub>2</sub>O to remove any ketones and allow it to stand overnight in the Et<sub>2</sub>O, then filter and dry it in air then in a vacuum. [Gomberg & Snow *J Am Chem Soc* 47 202 1925, Baines & Driver *J Chem Soc* 123 1216 1923, UV: Burawoy *Chem Ber* 64 462 1941, Beilstein 8 IV 2646.]

**Salicylaldehyde (*o*-hydroxybenzaldehyde)** [90-02-8] **M 122.1, b 93°/25mm, 195-197°/760mm, d<sub>4</sub><sup>20</sup> 1.167, n 1.574, pK<sup>25</sup> 8.37.** It is precipitated as the bisulfite addition compound by pouring the aldehyde slowly and with stirring into a 25% solution of NaHSO<sub>3</sub> in 30% EtOH, then standing for 30 minutes. The precipitate, after filtering at the pump, and washing with EtOH, is decomposed with aqueous 10% NaHCO<sub>3</sub>, and the aldehyde is extracted into diethyl ether, dried with Na<sub>2</sub>SO<sub>4</sub> or MgSO<sub>4</sub>, and distilled, under reduced pressure. *Alternatively*, salicylaldehyde is precipitated as its Cu complex by adding it to warm, saturated aqueous Cu(OAc)<sub>2</sub>, shaking and standing in ice. The precipitate is filtered off, washed with EtOH, then Et<sub>2</sub>O, and decomposed with 10% H<sub>2</sub>SO<sub>4</sub>; the aldehyde is extracted into Et<sub>2</sub>O, dried and vacuum distilled. It was also purified by dry column chromatography on Kieselgel G [Nishiya et al. *J Am Chem Soc* 108 3880 1986]. The *acetyl* derivative has **m 38-39°** (from petroleum ether or EtOH) and **b 142°/18mm, 253°/atm.** The *oxime*, [94-67-7] **M 137.1**, crystallises CHCl<sub>3</sub>/petroleum ether (b 40-60°) with **m 57°**. [Beilstein 8 IV 176, 203.]

**Salicylamide** [65-45-2] **M 137.1, m 142-144°, pK<sup>20</sup> 8.37.** Crystallise the amide from water or repeatedly from CHCl<sub>3</sub> [Nishiya et al. *J Am Chem Soc* 108 3880 1986]. [Beilstein 10 IV 169.] The *anilide* [87-17-2] **M 213.2, m 135°** crystallises from H<sub>2</sub>O. [Beilstein 12 H 500, 12 I 268, 12 II 256, 12 944.]

**Salicylhydroxamic acid** [89-73-61] **M 153.1, m 179-180°(dec), pK<sub>1</sub><sup>30</sup> 2.15, pK<sub>2</sub><sup>30</sup> 7.46, pK<sub>3</sub><sup>30</sup> 9.72.** Crystallise the hydroxamic acid from acetic acid. [Beilstein 10 H 98.]

**Salicylic acid (2-hydroxybenzoic acid)** [69-72-7] **M 138.1, m 157-159°, 158-160°, 159.5°, 159-160°, 162°, b 211°/20mm, pK<sub>1</sub><sup>25</sup> 3.01, pK<sub>2</sub><sup>25</sup> 13.43 (13.01).** It has been purified by steam distillation, by recrystallisation from H<sub>2</sub>O (solubility is 0.22% at room temperature and 6.7% at 100°), absolute MeOH, or cyclohexane and by sublimation in a vacuum at 76°. The *acid chloride* (needles) has **m 19-19.5°, b 92°/15mm**, the *amide* has **m 133°** (yellow needles from H<sub>2</sub>O), the *O-acetyl* derivative has **m 135°** (rapid heating and the liquid resolidifies at 118°), and the *O-benzoyl* derivative has **m 132°** (aqueous EtOH). [IR: Hales et al. *J Chem Soc* 3145 1954, Bergmann et al. *J Chem Soc* 2351 1950]. [Beilstein 10 IV 125.]

**Solochrome Violet R [4-hydroxy-3-(2-hydroxynaphthyl-1-ylazo)benzenesulfonic acid]** [2092-55-9] **M 367.3, CI 15670, λ<sub>max</sub> 501nm, pK<sub>2</sub><sup>25</sup> 7.22 (OH), pK<sub>3</sub><sup>25</sup> 13.39 (OH).** Convert the acid to the monosodium salt by precipitation with NaOAc/AcOH buffer of pH 4, then purify by precipitation of the free acid from aqueous solution with conc HCl, washing and extracting with EtOH in a Soxhlet extractor. The acid precipitates on evaporating the EtOH and is reconverted to the sodium salt as described for *Chlorazole Sky Blue FF*. Dry it at 110°. It is *hygroscopic*. [Coates & Rigg *Trans Faraday Soc* 57 1088 1961, Beilstein 16 II 127.]

**cis-Stilbene** [645-49-8] **M 180.3, b 145°/12mm.** Purify it by chromatography on alumina using hexane and distil it under vacuum. (The final product contains *ca* 0.1% of the *trans*-isomer.) [Lewis et al. *J Am Chem Soc* 107 203 1985, Saltiel *J Phys Chem* 91 2755 1987, Beilstein 5 H 630.]

**trans-Stilbene (trans-1,2-diphenylethylene)** [103-30-0] **M 180.3, m 125.9°, b 305-307°/744mm, d<sub>4</sub><sup>20</sup> 0.970.** Purify it by vacuum distillation. (The final product contains about 1% of the *cis* isomer.) Crystallise it from EtOH. It has also been purified by zone melting. The *stypnate* (see next entry) has **m 142°**. [Lewis et al. *J Am Chem Soc* 107 203 1985, Bollucci et al. *J Am Chem Soc* 109 515 1987, Saltiel *J Phys Chem* 91 2755 1987, Beilstein 5 IV 2156.]

**Styphnic acid (2,4,6-trinitroresorcinol)** [82-71-3] **M 245.1, m 177-178°, 179-180°, pK<sub>1</sub><sup>25</sup> 0.06 (1.74), pK<sub>2</sub><sup>25</sup> 4.23 (4.86).** Crystallise the phenol from ethyl acetate or water containing HCl [EXPLODES violently on rapid heating.] Its solubility in H<sub>2</sub>O is 0.7% at 20° and 3% at 100°. It forms *addition compounds* with aromatic hydrocarbons, e.g. naphthalene (**m 168°**), anthracene (**m 180°**), phenanthrene (**m 142°**), fluorene (**m 134°**) and retene (**m 141°**). [Beilstein 6 H 830, 6 III 4354, 6 IV 5699.]

**Styrene (vinylbenzene)** [100-42-5] **M 104.2, b 41-42°/18mm, 145.2°/760mm, d<sub>4</sub><sup>20</sup> 0.907, n<sub>D</sub><sup>20</sup> 1.5469, n<sub>D</sub><sup>25</sup> 1.5441.** Styrene is difficult to purify and keep pure. Usually it contains added inhibitors (such as a trace of hydroquinone). Wash it with aqueous NaOH to remove inhibitors (e.g. *tert*-butanol), then with water, dry it for several hours with MgSO<sub>4</sub> and distil it at 25° under reduced pressure in the presence of an inhibitor (such as 0.005% *p-tert*-butylcatechol). It can be stored at -78°. It can also be stored and kept anhydrous with Linde type 5A molecular sieves, CaH<sub>2</sub>, CaSO<sub>4</sub>, BaO or sodium, being fractionally distilled, and distilled in a vacuum line just before use. *Alternatively*, styrene (and its deuterated derivative) are passed through a neutral alumina column before use [Woon et al. *J Am Chem Soc* 108 7990 1986, Collman *J Am Chem Soc* 108 2588 1986]. [Beilstein 5 IV 1334.]

**(±)-Styrene glycol (±-1-phenyl-1,2-ethanediol)** [93-56-1] **M 138.2, m 67-68°, 272-274°/755mm.** Crystallise the diol from petroleum ether, Et<sub>2</sub>O, Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> (**m 69-70°**) or \*C<sub>6</sub>H<sub>6</sub>. The *dibenzoyl derivative* has **m 96-97°**. [Beilstein 6 H 907, 6 I 444, 6 II 887, 6 III 4572, 6 IV 5939.]

**Styrene oxide** [96-09-3] **M 120.2, b 84-86°/16.5mm, d<sub>4</sub><sup>20</sup> 1.053, n<sub>D</sub><sup>25</sup> 1.535.** Fractional distillation under reduced pressure does not remove phenylacetaldehyde. If this material is present, the styrene oxide is treated with hydrogen under 3 atmospheres pressure in the presence of platinum oxide. The aldehyde, but not the oxide, is reduced to β-phenylethanol, and separation is now readily achieved by fractional distillation. [Schenck & Kaizermen *J Am Chem Soc* 75 1636 1953, Beilstein 17/I V 577.]

**Sudan II (Sudan Blue, Solvent Blue 35, 1,4-bis-(butylamino)-9,10-anthraquinone)** [17354-14-2] **M 350.5, m 122°**,  $\lambda_{\max}$  604, 652nm,  $\text{pK}_{\text{Est}} \sim 9.5$  (OH). It is formed from quinizarin (2g see [81-64-1]), 33% EtOH/*n*-BuNH<sub>2</sub> (20ml) and Na<sub>2</sub>S<sub>2</sub>O<sub>4</sub> (2g) at 140°/8 hours, evaporate, extract with toluene, chromatograph (Al<sub>2</sub>O<sub>3</sub>), the intense blue band in toluene is evaporated, and the residue gave purple needles (Cu lustre) from petroleum ether (b 60-80°) (1.1g, 38%) [Peters & Walker *J Chem Soc* 1429 1956, *Beilstein* 14 IV 460]. It forms Cu and Ni salts.

**Sudan III [Solvent Red 23, 1-(*p*-phenylazo-phenylazo)-2-naphthol]** [85-86-9] **M 352.4, m 199°(dec), CI 26100**,  $\lambda_{\max}$  354, 508 nm,  $\text{pK}_{\text{Est}} \sim 9.0$ . Crystallise the dye from EtOH, EtOH/water or \*benzene/absolute EtOH (1:1). [*Beilstein* 16 II 75, 16 III 148, 16 IV 248.]

**Sudan IV [Solvent Red 24, 1-(4-*o*-tolylazo-*o*-tolylazo)-2-naphthol]** [85-83-6] **M 380.5, m ~184°(dec), CI 26105**,  $\lambda_{\max}$  520nm,  $\text{pK}_{\text{Est}} \sim 9.0$ . Crystallise the dye from EtOH/water or acetone/water. [*Beilstein* 16 IV 249.]

**Sulfaguanidine** [57-67-0] **M 214.2, m 189-190°**,  $\text{pK}_1$  0.48,  $\text{pK}_2$  2.75. Crystallise the antibacterial from hot water (7ml/g). [*Beilstein* 14 III 1970, 14 IV 2668.]

**Sulfanilic acid (4-aminobenzenesulfonic acid)** [121-57-3] **M 173.2,  $\text{pK}_1^{25} < 1$ ,  $\text{pK}_2^{25}$  3.23**. Crystallise the acid (as *dihydrate*) from boiling water. Dry it at 105° for 2-3 hours, then over 90% H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. The *S*-benzylisothiuronium salt has **m 187°** (from aqueous EtOH). [*Beilstein* 14 IV 2655.]

***o*-Sulfobenzoic acid (H<sub>2</sub>O)** [123333-68-6 (H<sub>2</sub>O), 632-25-7] **M 202.2, m 68-69°**,  $\text{pK}_{\text{Est}(1)} < 1$ ,  $\text{pK}_{\text{Est}(2)} \sim 3.1$  (CO<sub>2</sub>H). Crystallise the acid from water. The *S*-benzylisothiuronium salt has **m 205.5-206.5°** (from aqueous EtOH). [*Beilstein* 1 II 215, 1 III 658.] The *monoammonium salt* [6939-89-5] **M 219.5**, crystallises from water.

**5-Sulfosalicylic acid** [5965-83-3] **M 254.2, m 108-110°**,  $\text{pK}_1^{25} < 0$ ,  $\text{pK}_2^{25}$  2.67,  $\text{pK}_3^{25}$  11.67. Crystallise the acid from H<sub>2</sub>O. *Alternatively*, it is converted to the monosodium salt which is crystallised from H<sub>2</sub>O and washed with a little H<sub>2</sub>O, EtOH and then Et<sub>2</sub>O. The acid is recovered by acidifying. The *S*-4-chlorobenzylisothiuronium salt has **m 181°** (from dioxane). [*Beilstein* 11 H 411, 11 II 232, 11 III 704.]

**Syringaldehyde (3,5-dimethoxy-4-hydroxybenzaldehyde)** [134-96-3] **M 182.2, m 113°**,  $\text{pK}_{\text{Est}} \sim 8$ . Crystallise syringaldehyde from petroleum ether. [*Beilstein* 8 H 391, 8 IV 2718.]

**Syringic acid (3,5-dimethoxy-4-hydroxybenzoic acid)** [530-57-4] **M 198.2, m 204-205°, 206.5°, 206-209°, 209-210°**,  $\text{pK}_1^{25}$  4.34,  $\text{pK}_2^{25}$  9.49. Recrystallise syringic acid from H<sub>2</sub>O using charcoal [Bogert & Coyne *J Am Chem Soc* 51 571 1929, Anderson & Nabenhauer *J Am Chem Soc* 48 3001 1926.] The *methyl ester* has **m 107°** (from MeOH), the *4-acetyl* derivative has **m 190°** and the *4-benzoyl* derivative has **m 229-232°**. [Hahn & Wassmuth *Chem Ber* 67 2050 1934, UV: Lemon *J Am Chem Soc* 69 2998 1947 and Pearl & Beyer *J Am Chem Soc* 72 1743 1950, *Beilstein* 10 IV 1995.]

**Terephthalaldehyde** [623-27-8] **M 134.1, m 116°, b 245-248°/771mm**. Crystallise terephthalaldehyde from water. [*Beilstein* 7 IV 2140.]

**Terephthalic acid (benzene-1,4-dicarboxylic acid)** [100-21-0] **M 166.1, m sublimes >300° without melting**,  $\text{pK}_1^{20}$  3.4,  $\text{pK}_2^{20}$  4.34. Purify the acid *via* the sodium salt which, after crystallisation from water, is re-converted to the acid by acidification with mineral acid. Filter off the solid, wash it with H<sub>2</sub>O and dry it in a vacuum. The *S*-benzylisothiuronium salt has **m 204°** (from aqueous EtOH). [*Beilstein* 9 IV 3301.]

**Terephthaloyl chloride** [100-20-9] **M 203.0, m 80-82°**. Crystallise the acid chloride from dry hexane. [*Beilstein* 9 IV 3318.]

***o*-Terphenyl** [84-15-1] **M 230.3, m 58-59°, b 337°/atm**. Crystallise *o*-terphenyl from EtOH. Also purify it by chromatography of CCl<sub>4</sub> solution on alumina, with petroleum ether as eluent, followed by crystallisation from petroleum ether (b 40-60°) or petroleum ether/\*C<sub>6</sub>H<sub>6</sub>. It also distils under vacuum. [*Beilstein* 5 III 2292, 5 IV

2478.]

*m*-Terphenyl [92-06-8] M 230.3, m 88-89°, b 379°/atm. Purify it as for *o*-terphenyl above. [Beilstein 5 IV 2480.]

*p*-Terphenyl (1,4-diphenylbenzene) [92-94-4] M 230.3, m 212.7°, b 389°/atm. Crystallise *p*-terphenyl from nitrobenzene or trichlorobenzene. It is also purified by chromatography on alumina in a darkened room, using petroleum ether, and then crystallising from petroleum ether (b 40-60°) or petroleum ether/\*benzene. It is a fluorophore for scintillation counting and has  $\lambda_{\text{ex}}$  286nm :  $\lambda_{\text{em}}$  343nm in DMF, and  $\lambda_{\text{max}}$  at 277nm (log $\epsilon$  4.50). [Beilstein 5 IV 2483.]

Terreic acid (2-hydroxy-3-methyl-1,4-benzoquinone-5,6-epoxide) [121-40-4] M 154.1, m 127-127.5°, [ $\alpha$ ]<sub>D</sub><sup>22</sup> +74° (pH 4, phosphate buffer), -17° (CHCl<sub>3</sub>), pK<sup>25</sup> 4.5. Crystallise terreic acid from \*C<sub>6</sub>H<sub>6</sub>, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or hexane. Sublime it 80-100°/1mm. [Beilstein 17 IV 6698.]

3',3'',5',5''-Tetrabromophenolphthalein ethyl ester [1176-74-5] M 662.0, m 212-214°. Crystallise the ester from \*benzene, dry at 120° and keep it under vacuum. [Beilstein 10 III 4490.]

2,3,4,5-Tetrachloroaniline [634-83-3] M 230.9, m 117-118°, 119-120°, pK<sub>Est</sub> ~-0.26. Crystallise it from EtOH. The *acetyl* derivative has m 165-166° (from EtOH). [Beilstein 12 H 630, 12 I 313, 12 II 340, 12 IV 1286.]

2,3,5,6-Tetrachloroaniline [3481-20-7] M 230.9, m 107-108°, pK<sub>Est</sub> ~-1.8. Crystallise it from EtOH. The *acetyl* derivative has m 213-214° (from EtOH). [Beilstein 12 II 340, 12 III 1414, 12 IV 1287.]

1,2,3,4-Tetrachlorobenzene [634-66-2] M 215.9, m 45-46°, 47.5°, b 254°/760mm. Crystallise it from EtOH. [Beilstein 5 H 204, 5 II 156, 5 III 550, 5 IV 667.]

1,2,3,5-Tetrachlorobenzene [634-90-2] M 215.9, m 50-51°, 51°, 54-55°, b 246°/760mm, d<sup>10</sup> 1.7344. Crystallise it from EtOH. [Beilstein 5 II 157, 5 III 551, 5 IV 668.]

1,2,4,5-Tetrachlorobenzene [95-94-3] M 215.9, m 139.5-140.5°, b 240°/760mm. Crystallise it from EtOH, ether, \*benzene, \*benzene/EtOH or carbon disulfide. [Beilstein 5 IV 668.]

3,4,5,6-Tetrachloro-1,2-benzoquinone (*o*-chloranil) [2435-53-2] M 245.9, m 130°. Crystallise *o*-chloranil from AcOH. Dry it in vacuum desiccator over KOH. [Beilstein 7 IV 2068.]

Tetrachloro-*N*-methylphthalimide [14737-80-5] M 298.9, m 209.7°. Crystallise the imide from absolute EtOH. [Beilstein 21 H 505, 17/11 V 260.]

2,3,4,6-Tetrachloronitrobenzene (1,2,3,5-tetrachloro-4-nitrobenzene) [879-39-0] M 260.9, m 41-42°, 42°. Crystallise it from aqueous EtOH. [Beilstein 5 II 187, 5 III 617, 5 IV 728.]

2,3,5,6-Tetrachloronitrobenzene (1,2,4,5-tetrachloro-3-nitrobenzene) [117-18-0] M 260.9, m 99-100°. Crystallise it from aqueous EtOH or H<sub>2</sub>O. [Beilstein 5 III 617, 5 IV 728.]

2,3,4,5-Tetrachlorophenol [4901-51-3] M 231.9, m 116-117°, pK<sup>25</sup> 6.95. Crystallise the phenol from petroleum ether. The *benzoate* has m 110° (from EtOH). [Beilstein 6 II 182, 6 III 729, 6 IV 1020.]

2,3,4,6-Tetrachlorophenol [58-90-2] M 231.9, m 70°, b 150°/15mm, pK<sup>25</sup> 5.38. Crystallise the phenol from petroleum ether. The *benzoate* has m 116° (from EtOH). [Beilstein 6 H 193, 6 III 729, 6 IV 1021.]

2,3,5,6-Tetrachlorophenol [935-95-5] M 231.9, m 115°, pK<sup>25</sup> 5.48. Crystallise the phenol from petroleum ethers. The *benzoate* has m 136° (from EtOH). [Beilstein 6 II 182, 6 III 730, 6 IV 1025.]

**Tetrachlorophthalic anhydride** [117-08-8] **M 285.9, m 255-257°**. Crystallise the anhydride from chloroform or \*benzene, then sublime it in a vacuum. [Beilstein 17/11 V 260.]

**1,2,4,5-Tetracyanobenzene** [712-74-3] **M 178.1, m 270-272° (280°)**. Crystallise the tetra-nitrile from EtOH and sublime *in vacuo*. [Lawton & McRitchie *J Org Chem* 24 26 1959, Bailey et al. *Tetrahedron* 19 161 1963, *Beilstein* 9 IV 3800.]

**Tetrahydroxy-*p*-benzoquinone (2H<sub>2</sub>O)** [5676-48-2; 123334-16-7 2H<sub>2</sub>O] **M 208.1, pK<sub>1</sub><sup>30</sup> 4.80, pK<sub>2</sub><sup>30</sup> 6.8**. Crystallise the quinone from water. [Beilstein 8 H 534, 8 II 572, 8 III 4204, 8 IV 3604.]

**Tetralin (1,2,3,4-tetrahydronaphthalene)** [119-64-2] **M 132.2, m -35.79° (from CF<sub>2</sub>Cl<sub>2</sub>), b 65-66°/5mm, 207.6°/760mm, d<sub>4</sub><sup>20</sup> 0.968, n<sub>D</sub><sup>20</sup> 1.5413**. Wash tetralin with successive portions of conc H<sub>2</sub>SO<sub>4</sub> until the acid layer is no longer coloured, then wash it with aqueous 10% Na<sub>2</sub>CO<sub>3</sub>, and then distilled water. Dry (CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>), filter, reflux and fractionally distil it under reduced pressure from sodium or BaO. It can also be purified by repeated fractional freezing.

Bass [*J Chem Soc* 3498 1964] freed tetralin, purified as above, from naphthalene and other impurities by conversion to ammonium tetralin-6-sulfonate. Concentrated H<sub>2</sub>SO<sub>4</sub> (150ml) is added slowly to stirred tetralin (272ml) which is then heated on a water bath for about 2 hours for complete solution. The warm mixture, when poured into aqueous NH<sub>4</sub>Cl solution (120g in 400ml water), gives a white precipitate which, after filtering off, is crystallised from boiling water, washed with 50% aqueous EtOH and dried at 100°. Evaporation of its boiling aqueous solution on a steam bath removes traces of naphthalene. The pure salt (229g) is mixed with conc H<sub>2</sub>SO<sub>4</sub> (266ml) and steam distilled from an oil bath at 165-170°. An ether extract of the distillate is washed with aqueous Na<sub>2</sub>SO<sub>4</sub>, and the ether is evaporated, prior to distilling the tetralin from sodium. Tetralin has also been purified *via* barium tetralin-6-sulfonate, conversion to the sodium salt and decomposed in 60% H<sub>2</sub>SO<sub>4</sub> using superheated steam. [Beilstein 5 H 491, 5 III 1219, 5 IV 1388.]

**Tetralin hydroperoxide** [771-29-9] **M 164.2, m 55.7-56°, 56°**. Crystallise the tetralin hydroperoxide from hexane, toluene at -30° (m 54.0-54.5°). The oxygen content should be ~9.70%. [Knight & Swern *Org Synth Coll Vol* 1V 895 1963.]

**α-Tetralone (1,2,3,4-tetrahydro-1-oxonaphthalene)** [529-34-0] **M 146.2, m 2-7°, 7.8-8.0°, b 75-85°/0.3mm, 89°/0.5mm, 94-95°/2mm, 132-134°/15mm, 143-145°/20mm, d<sub>4</sub><sup>20</sup> 1.0695, n<sub>D</sub><sup>20</sup> 1.5665**. Check the IR first. Purify α-tetralone by dissolving 20ml in Et<sub>2</sub>O (200ml), washing with H<sub>2</sub>O (100ml), 5% aqueous NaOH (100ml), H<sub>2</sub>O (100ml), 3% aqueous AcOH (100ml), 5% NaHCO<sub>3</sub> (100ml) then H<sub>2</sub>O (100ml) and dry the ethereal layer over MgSO<sub>4</sub>. Filter, evaporate and fractionate the residue through a 6in Vigreux column under reduced pressure to give a colourless oil (~17g) with **b** 90-91°/0.5-0.7mm. [Snyder & Werber *Org Synth Coll Vol* III 798 1955.] It has also been fractionated through a 0.5metre packed column with a heated jacket under reflux using a partial take-off head. It has λ<sub>max</sub> at 247.5 and 290nm (hexane). The *phenylhydrazone* has **m** 83°. The *2,4,6-trinitrophenylhydrazone* has **m** 247.5-248° (from EtOH). [Olson & Bader *Org Synth Coll Vol* IV 898 1963, *Beilstein* 7 H 370, 7 III 1416, 7 IV 1015.]

**β-Tetralone (1,2,3,4-tetrahydro-2-oxonaphthalene)** [530-93-8] **M 146.2, m 17-18°, ~18°, b 93-95°/2mm, 104-105°/4mm, 114-115°/4-5mm, 140°/18mm, d<sub>4</sub><sup>20</sup> 1.1000, n<sub>D</sub><sup>20</sup> 1.5598**. If reasonably pure, then fractionate it through an efficient column. Otherwise purify it *via* the *bisulfite adduct*. To a solution of NaHSO<sub>3</sub> (32.5g, 0.31mol) in H<sub>2</sub>O (57ml) is added 95% EtOH (18ml) and set aside overnight. Any bisulfite-sulfate that separated is removed by filtration, and the filtrate is added to the tetralone (14.6g, 0.1mol) and shaken vigorously. The adduct separates in a few minutes as a white precipitate and is kept on ice for ~3.5 hours with occasional shaking. The precipitate is collected, washed with 95% EtOH (13ml), then with Et<sub>2</sub>O (4 × 15ml, by stirring the suspension in the solvent, filtering and repeating the process). The colourless product is dried in air and stored in air tight containers in which it is stable for extended periods (yield is ~17g). This bisulfite (5g) is suspended in H<sub>2</sub>O (25ml), and Na<sub>2</sub>CO<sub>3</sub>·H<sub>2</sub>O (7.5g) is added (pH of solution is ~10). The mixture is then extracted with Et<sub>2</sub>O (5 × 10ml, i.e. until the aqueous phase does not test for tetralone — see below). Wash the combined extracts with 10% aqueous HCl (10ml), H<sub>2</sub>O (10ml, i.e. until the washings are neutral), dry (MgSO<sub>4</sub>), filter, evaporate and distil the residual oil using a Claisen flask under reduced pressure and in a N<sub>2</sub> atmosphere. The pure

tetralone is a colourless liquid **b** 70-71°/0.25mm (see also above). The yield is ~2g. **Tetralone test:** Dissolve a few drops of the tetralone solution (ethereal or aqueous) in 95% EtOH in a test tube and add 10 drops of 25% NaOH down the side of the tube. A deep blue colour develops at the interface with air. [Soffer et al. *Org Synth Coll Vol IV* 903 1963, Cornforth et al. *J Chem Soc* 689 1942, UV: Soffer et al. *J Am Chem Soc* 1556 1952.] The *phenylhydrazone* has **m** 108° [Crawley & Robinson *J Chem Soc* 2001 1938]. [*Beilstein* 7 H 370, 7 II 295, 7 III 1422, 7 IV 1018.]

**1,2,3,4-Tetramethylbenzene (prehnitine)** [488-23-3] **M 134.2, m -6.3°, b 79.4°/10mm, 204-205°/760mm,  $d_4^{20}$  0.905,  $n_D^{20}$  1.5203.** Dry it over sodium and distil under reduced pressure. The *picrate* has **m** 92-95° (EtOH). [*Beilstein* 5 H 430, 5 I 206, 5 II 329, 5 III 974, 5 IV 1072.]

**1,2,3,5-Tetramethylbenzene (isodurene)** [527-53-7] **M 134.2, m -23.7°, b 74.4°/10mm, 198°/760mm,  $d_4^{20}$  0.890,  $n_D^{20}$  1.5130.** Reflux isodurene over sodium and distil it under reduced pressure. [Smith *Org Synth Coll Vol II* 248 1943, *Beilstein* 5 H 430, 5 II 329, 5 III 976, 5 IV 1073.]

**1,2,4,5-tetramethylbenzene (durene)** [95-93-2] **M 134.2, m 79.5-80.5°, 191-192°/760mm.** Chromatograph durene on alumina, and recrystallise it from aqueous EtOH or \*benzene. Zone-refining removes duroaldehydes. Dry it under vacuum. [Yamauchi et al. *J Phys Chem* 89 4804 1985.] It has also been sublimed *in vacuo* [Johnston et al. *J Am Chem Soc* 109 1291 1987]. [*Beilstein* 5 H 431, 5 I 207, 5 II 329, 5 III 979, 5 IV 1076.]

***N,N,N',N'*-Tetramethylbenzidine** [366-29-0] **M 240.4, m 195.4-195.6°,  $pK_{Est(1)}$  ~3.4,  $pK_{Est(2)}$  ~4.5.** Crystallise the benzidine from EtOH or petroleum ether, then from petroleum ether/\*benzene, and sublime it in a vacuum. [Guarr et al. *J Am Chem Soc* 107 5104 1985.] Dry it *in vacuo* in a drying pistol, or a vacuum line. It has **m** 195-196° after sublimation. [*Beilstein* 13 H 221, 13 I 61, 13 II 97, 13 III 429, 13 IV 368.]

***p,p'*-Tetramethyldiaminodiphenylmethane [bis(*p*-dimethylaminophenyl)methane, Michler's base, (*p,p'*-methylene-bis-*N,N*-dimethylaniline)]** [101-61-1] **M 254.4, m 89-90°, b 155-157°/0.1mm,  $pK_{Est(1)}$  ~5.8,  $pK_{Est(2)}$  ~5.1.** Crystallise the base from EtOH (2ml/g) or 95% EtOH (*ca* 12ml/g). It sublimes on heating. [*Beilstein* 13 IV 390.]

***N,N,N',N'*-Tetramethyl-1,8-naphthalenediamine [Proton sponge, 1,8-bis-(dimethylamino)-naphthalene** [20734-58-1] **M 214.3, m 45-48°, 47-48°, b 144-145°/4mm,  $pK_1$  -10.5 (from half protonation in 86% aqueous H<sub>2</sub>SO<sub>4</sub>, diprotonation),  $pK_2$  12.34 (monoprotonation).** It is prepared by methylating 1,8-diaminonaphthalene, and likely impurities are methylated products. The tetramethyl compound is a stronger base than the unmethylated, di and trimethylated derivatives. The  $pK_a$  values are: 1,8-(NH<sub>2</sub>)<sub>2</sub> = 4.61, 1,8-(NHMe)<sub>2</sub> = 5.61, 1-NHMe-8-NHMe<sub>2</sub> = 6.43 and 1,8-(NMe<sub>2</sub>)<sub>2</sub> = 12.34. The mixture is then treated with H<sub>2</sub>O at pH 8 (where all but the required base are protonated) and extracted with Et<sub>2</sub>O or CHCl<sub>3</sub>. The dried extract (K<sub>2</sub>CO<sub>3</sub>) yields the tetramethyldiamine on evaporation which can be distilled. It is a strong base with weak nucleophilic properties, e.g. it could not be alkylated by refluxing with EtI in MeCN for 4 days; and on treatment with methyl fluorosulfonate only the fluorosulfonate salt of the base is obtained. [NMR: Adler et al. *J Chem Soc, Chem Commun* 723 1968, Brown & Letang *J Am Chem Soc* 63 358 1941, Brzezinski et al. *J Chem Soc Perkin Trans 2* 857 1991.] Alternatively, crystallise proton sponge from EtOH and dry it in a vacuum oven. Store it in the dark in a CO<sub>2</sub>-free atmosphere. [Benoit et al. *Can J Chem* 65 996 1987, *Beilstein* 13 IV 344.]

***N,N,N',N'*-Tetramethyl-1,4-phenylenediamine** [100-22-1] **M 164.3, m 51°, b 135°/14mm, 260°/760mm,  $pK_1^{20}$  2.29,  $pK_2^{20}$  6.35.** Crystallise the amine from petroleum ether or water. It can be sublimed or dried carefully in a vacuum line, and stored in the dark under nitrogen. It has been recrystallised from its melt. [*Beilstein* 13 H 74, 13 I 22, 13 II 40, 13 III 111, IV 107.]

***N,N,N',N'*-Tetramethyl-1,4-phenylenediamine dihydrochloride (Wurster's Reagent)** [637-01-4] **M 237.2, m 222-224°.** Crystallise the salt from isopropyl or *n*-butyl alcohols, saturated with HCl. Treat it with aqueous NaOH to give the *free base* (see previous entry) which is filtered, dried and sublimed in a vacuum. [Guarr et al. *J Am Chem Soc* 107 5104 1985, *Beilstein* 13 H 74.] Oxidase reagent (1% aqueous solution) used for testing cytochrome +ve aerobic microorganisms.

**Tetra(*p*-nitrophenyl)ethylene** [47797-98-8] **M 512.4, m 298-299°**. Crystallise it from dioxane or AcOH (**m** 292°, yellow needles), and dry it at 150°/0.1mm. [Gorvin *J Chem Soc* 678 1959, Schlenk *Justus Liebigs Ann Chem* 394 214 1913, *Beilstein* 5 H 744, 5 IV 2780.]

**Tetraphenylethylene** [632-51-9] **M 332.4, m 223-224°, b 415-425°/760mm**. Crystallise the ethylene from dioxane or from EtOH/\*C<sub>6</sub>H<sub>6</sub>. Sublime it under high vacuum. [*Beilstein* 5 H 743, 5 IV 2780.]

**Tetraphenylhydrazine** [632-52-0] **M 336.4, m 147°, pK<sub>Est</sub> ~0**. Crystallise the hydrazine from 1:1 CHCl<sub>3</sub>/toluene, 1:5 CHCl<sub>3</sub>/EtOH (**m** 149°), \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. Store it in a refrigerator, in the dark. [*Beilstein* 15 H 125, 15 I 29, 15 III 77, 15 IV 59.]

**trans-1,1,4,4-Tetraphenyl-2-methylbutadiene** [20411-57-8] **M 372.5**. Crystallise it from EtOH.

**5,6,11,12-Tetraphenylanthracene (Rubrene)** [517-51-1] **M 532.7, m>315°, 322°, d<sub>4</sub><sup>20</sup> 1.255**. Rubrene forms orange crystals on sublimation at 250-260°/3-4mm [UV Badger & Pearce *Spectrochim Acta* 4 280 1950]. It has also been recrystallised from \*benzene under red light because it is chemiluminescent and light sensitive. [*Beilstein* 5 IV 2968.]

**1,2,3,4-Tetraphenylanthralene** [751-38-2] **M 432.6, m 199-201°, 204-204.5°**. Crystallise the anthralene from MeOH or EtOH. [Fieser & Haddadin *Org Synth* 46 107 1966, *Beilstein* 5 IV 2918.]

**Thioacetanilide** [677-53-6] **M 151.2, m 75-76°, 76-79°, pK<sub>Est</sub> ~13.1**. Crystallise thioacetanilide from H<sub>2</sub>O and dry it *in vacuo*. [*Beilstein* 12 H 245, 12 I 193, 12 II 142, 12 III 464, 12 IV 378.]

**Thiobenzanilide** [636-04-4] **M 213.2, m 101.5-102°, pK<sub>Est</sub> ~12.6**. Crystallise thiobenzanilide from MeOH at Dry-ice temperature.

**Thio-Michler's Ketone [4,4'-bis(dimethylamino)thiobenzophenone]** [1226-46-6] **M 284.4, m 102-106°, λ<sub>max</sub> 457 nm (ε 2.92 × 10<sup>4</sup> in 30% aqueous *n*-propanol)**. Purify the thioketone by recrystallisation from hot EtOH or by trituration with a small volume of CHCl<sub>3</sub>, followed by filtration and washing with hot EtOH [Terbell & Wystrade *J Phys Chem* 68 2110 1964]. Also it recrystallises from CHCl<sub>3</sub>/ MeOH. [*Beilstein* 14 H 101, 14 I 395, 14 II 60.]

**1-Thionaphthol** [529-36-2] **M 160.2, b 106°/1.5mm, 208.5°/200mm, d<sub>4</sub><sup>20</sup> 1.161, n<sub>D</sub><sup>20</sup> 1.6802, pK<sup>25</sup> 6.34**. It is steam volatile and is purified by distillation in the absence of O<sub>2</sub>, as it oxidises to the disulfide. It is soluble in Et<sub>2</sub>O and EtOH but very slightly soluble in H<sub>2</sub>O and dilute alkalis. The *S*-ethyl derivative [17539-31-0] **M 188.2**, has **b** 175-176°/25mm, and **d**<sup>0</sup> 1.120. [*Beilstein* 6 III 2943, 6 IV 4241.]

**2-Thionaphthol** [91-60-1] **M 160.2, m 81.8-82.4°, 82°, b 153.5°/15mm, 286°/760mm, pK<sup>25</sup> 6.47**. It is steam volatile. It has to be distilled under Ar or N<sub>2</sub>, as it oxidises to the disulfide, and crystallises from EtOH. The *S*-methyl derivative has **m** 104-105° (from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether), and the *S*-ethyl derivative [32551-87-4] **M 188.2**, has **m** 16° and **b** 175-170.5°/15mm. The *S*-acetate [831-23-2] has **m** 53.5° and **b** 191°/15mm, and the *diethylamine salt* forms yellow needles **m** 107° from dioxane. [*Beilstein* 6 H 657, 6 I 316, 6 II 610, 6 III 3006, 6 IV 4312.]

**Thiophenol (benzenethiol)** [108-98-5] **M 110.2, f -14.9°, b 46.4°/10mm, 168.0°/760mm, d<sub>4</sub><sup>20</sup> 1.073, n<sub>D</sub><sup>20</sup> 1.5897, pK<sup>25</sup> 6.62**. Dry thiophenol with CaCl<sub>2</sub> or CaSO<sub>4</sub>, and distil it at 10mm pressure or at 100mm (**b** 103.5°) in a stream of nitrogen. The *2,4*-dinitrophenyl thioether has **m** 121° (from EtOH), and the *2,4*-dinitrophenyl sulfone has **m** 161° (from EtOH). [*Beilstein* 6 IV 1463.]

**Thiosalicylic (2-mercaptobenzoic) acid** [147-93-3] **M 154.2, m 164-165°, pK<sub>1</sub><sup>25</sup> 3.54, pK<sub>2</sub><sup>25</sup> 8.80**. Crystallise the thio acid from hot EtOH (4ml/g), after adding hot distilled water (8ml/g) and boiling with charcoal. The hot solution is filtered, cooled, the solid is collected and dried *in vacuo* (P<sub>2</sub>O<sub>5</sub>). Crystallise it from AcOH and sublime *in vacuo*. [*Beilstein* 10 IV 272.]

**Thymol (2-isopropyl-5-methylphenol)** [89-83-8] **M 150.2, m 49-51°, 51.5°, b 232°/atm,  $d_4^{25}$  0.965,  $n_D^{25}$  1.5204,  $pK^{20}$  10.62.** It occurs in the volatile oils of *Thymus vulgaris* and *Moranda punctata* L. from which it was first isolated [Arppe *Justus Liebigs Ann Chem* **58** 41 1846]. It is quite volatile at 100°, and is steam volatile, separating as white crystals which have a characteristic pungent odour, with a caustic taste. It should be stored in the dark, preferably under N<sub>2</sub> as it oxidises slowly in air. It possesses antimould properties, is an internal and external antiseptic, and has nematode anthelmintic activity. The *acetate* [528-93-2] is a yellow irritating oil, **b 243.5-245.5°/atm,  $d^0$  1.009,** with the odour of thymol, and is soluble in most organic solvents, but almost insoluble in H<sub>2</sub>O. [Beilstein **6** IV 3334.]

**Thymolphthalein complexone** [1913-93-5] **M 720.8, m 190°(dec),  $pK_1^{18.2}$  7.35,  $pK_2^{18.2}$  12.25.** Purify it as for phthalein complexone except that it is synthesised from thymolphthalein instead of cresolphthalein. [Beilstein **18/4** V 194.]

***o*-Tolidine (3,3'-dimethylbenzidine)** [119-93-7] **M 212.3, m 131-132°,  $pK^{25}$  4.45.** Dissolve the tolidine in \*benzene by percolation through a column of activated alumina and crystallise it from \*benzene/petroleum ether. [Beilstein **13** IV 410.]

***p*-Tolualdehyde** [104-87-0] **M 120.2, b 83-85°/0.1mm, 199-200°/760mm,  $d_4^{20}$  1.018,  $n_D^{20}$  1.548.** Steam distil the aldehyde, dry it with CaSO<sub>4</sub>, then fractionally distil it. [Beilstein **7** IV 672.]

***o*-Toluamide** [527-85-5] **M 135.2, m 141°, 144-145°.** Crystallise *o*-toluamide from hot water (10ml/g) and dry in air. [Noller *Org Synth Coll Vol II* 586 1943, Beilstein **9** H 465, **9** II 319, **9** III 2304.]

**Toluene** [108-88-3] **M 92.1, b 110.6°,  $d_4^{10}$  0.87615,  $d_4^{25}$  0.86231,  $n_D^{20}$  1.49693,  $n_D^{25}$  1.49413.** Dry toluene with CaCl<sub>2</sub>, CaH<sub>2</sub> or CaSO<sub>4</sub>, and dry further by standing with sodium, P<sub>2</sub>O<sub>5</sub> or CaH<sub>2</sub>. It can be fractionally distilled from sodium or P<sub>2</sub>O<sub>5</sub>. Unless specially purified, toluene is likely to be contaminated with methylthiophenes and other sulfur-containing impurities. These can be removed by shaking with conc H<sub>2</sub>SO<sub>4</sub>, but the temperature must be kept below 30° if sulfonation of toluene is to be avoided. A typical procedure consists of shaking toluene twice with cold conc H<sub>2</sub>SO<sub>4</sub> (100ml of acid per L), once with water, once with aqueous 5% NaHCO<sub>3</sub> or NaOH, again with H<sub>2</sub>O, then drying successively with CaSO<sub>4</sub> and P<sub>2</sub>O<sub>5</sub>, with final distillation from P<sub>2</sub>O<sub>5</sub> or over LiAlH<sub>4</sub> after refluxing for 30 minutes. *Alternatively*, the treatment with NaHCO<sub>3</sub> can be replaced by boiling under reflux with 1% sodium amalgam. Sulfur compounds can also be removed by prolonged shaking of the toluene with mercury, or by two distillations from AlCl<sub>3</sub>, the distillate then being washed with water, dried with K<sub>2</sub>CO<sub>3</sub> and stored with sodium wire. Other purification procedures include refluxing and distillation of sodium dried toluene from diphenylpicrylhydrazyl, and from SnCl<sub>2</sub> (to ensure freedom from peroxides). It has also been co-distilled with 10% by volume of ethyl methyl ketone, and again fractionally distilled. [Brown & Pearsall *J Am Chem Soc* **74** 191 1952.] For removal of carbonyl impurities see \*benzene. Toluene has been purified by distillation under nitrogen in the presence of sodium benzophenone ketyl. Toluene has also been dried with MgSO<sub>4</sub>, after the sulfur impurities have been removed, and then fractionally distilled from P<sub>2</sub>O<sub>5</sub> and stored in the dark [Tabushi et al. *J Am Chem Soc* **107** 4465 1985]. Toluene can be purified by passage through a tightly packed column of Fuller's earth.

**Rapid purification:** Alumina, CaH<sub>2</sub> and 4A molecular sieves (3% w/v) may be used to dry toluene (6 hours stirring and standing). Then the toluene is distilled, discarding the first 5% of distillate, and is stored over molecular sieves (3A, 4A) or Na wire. [Beilstein **5** H 280, **5** I 144, **5** II 209, **5** III 651, **5** IV 766.]

**Toluene-2,4-diamine (4-methyl-*m*-phenylenediamine)** [95-80-7] **M 122.2, m 99°, b 148-150°/8mm, 292°/760mm,  $pK_{Est(1)}$  ~2.5,  $pK_{Est(2)}$  ~4.4.** Recrystallise the diamine from water containing a very small amount of sodium dithionite (to prevent air oxidation), and dry it under vacuum. It also crystallises from \*benzene. [Beilstein **13** IV 235.]

***o*-Toluenesulfonamide** [88-19-7] **M 171.2, m 155.5°.** Crystallise the amide from hot H<sub>2</sub>O (**m** 153°), then from EtOH or Et<sub>2</sub>O/petroleum ether. The *N*-*o*-toluenesulfonylphthalimide has **m** 182° (from EtOH). [Evans & Dehn *J Am Chem Soc* **51** 3652 1929, Beilstein **11** H 86, **11** I 23, **11** II 39, **11** III 167, **11** IV 229.]

***p*-Toluenesulfonamide** [70-55-3] **M 171.2, m 137-137.5°, 138°**. Crystallise the amide from hot water, then from EtOH or Et<sub>2</sub>O/petroleum ether. [Beilstein 11 H 104, 11 IV 376.]

***p*-Toluenesulfonic acid** [6192-52-5] **M 190.2, m 38° (anhydrous), m 105-107° (monohydrate), pK<sup>25</sup> 1.55**. Purify the acid by precipitation from a saturated solution at 0° by introducing HCl gas. It can also be crystallised from conc HCl, then crystallised from dilute HCl (charcoal) to remove benzenesulfonic acid. It has been crystallised from EtOH/water. Dry it in a vacuum desiccator over solid KOH and CaCl<sub>2</sub>. *p*-Toluenesulfonic acid can be dehydrated by azeotropic distillation with \*benzene or by heating at 100° for 4 hours under water-pump vacuum. The anhydrous acid can be crystallised from \*benzene, CHCl<sub>3</sub>, ethyl acetate, anhydrous MeOH, or from acetone by adding a large excess of \*benzene. It can also be dried under vacuum at 50°. The *S*-benzylisothiuronium salt has **m 182°** (from aqueous EtOH). [Beilstein 11 IV 241.]

**Toluenesulfonic acid hydrazide (tosylhydrazide)** [1576-35-8] **M 186.2, m 108-110°, 109-110°**. Dissolve the hydrazide in hot MeOH (~1g/4ml), filter through Celite and precipitate the material by adding 2-2.5 volumes of distilled H<sub>2</sub>O. Dry it in air or in a vacuum. [Fiedman et al. *Org Synth Coll Vol V 1055 1973, Beilstein 11 II 66.*]

***p*-Toluenesulfonyl chloride (tosyl chloride)** [98-59-9] **M 190.7, m 66-69°, 67.5-68.5°, 69°, b 138-139°/9mm, 146°/15mm, 167°/36mm**. Material that has been standing for a long time contains tosic acid and HCl and has **m ca 65-68°**. It is purified by dissolving (10g) in the minimum volume of CHCl<sub>3</sub> (ca 25ml) filtered, and diluted with five volumes (i.e. 125ml) of petroleum ether (b 30-60°) to precipitate impurities. The solution is filtered, clarified with charcoal and concentrated to 40ml by evaporation. Further evaporation to a very small volume gives 7g of white crystals which are analytically pure, **m 67.5-68.5°**. (The insoluble material is largely tosic acid and has **m 101-104°**.) [Pelletier *Chem Ind (London) 1034 1953.*]

It also crystallises from toluene/petroleum ether in the cold, from petroleum ether (b 40-60°) or \*benzene. Its solution in diethyl ether has been washed with aqueous 10% NaOH until colourless, then dried (Na<sub>2</sub>SO<sub>4</sub>) and crystallised by cooling in powdered Dry-ice. It has also been purified by dissolving in \*benzene, washing with aqueous 5% NaOH, then dried with K<sub>2</sub>CO<sub>3</sub> or MgSO<sub>4</sub>, and distilled under reduced pressure and can be sublimed at high vacuum [Ebel *Chem Ber 60 2086 1927.*] [Beilstein 11 IV 375.]

***p*-Toluenethiol (*p*-thiocresol)** [106-45-6] **M 124.2, m 43.5-44°, pK<sup>25</sup> 6.82**. Crystallise the thiol from petroleum ether (b 40-70°). The *2,4*-dinitrophenyl thioether has **m 103°** (from EtOH), and the *2,4*-dinitrophenyl sulfone has **m 190°** (from EtOH). [Beilstein 6 IV 2153.]

***p*-Toluenethiosulfonic acid potassium salt (potassium *p*-toluenethiosulfonate)** [28519-50-8] **M 226.4, m 227-229°**. When a solution of KOH (64.9g, 86.5%, 1.0 mole) in H<sub>2</sub>O (28ml) is cooled in an ice bath while it is saturated with H<sub>2</sub>S (in an efficient **fume cupboard**) and excess of H<sub>2</sub>S is flushed out with N<sub>2</sub> (important to remove excess of H<sub>2</sub>S), a fresh solution of KHS is obtained. This is diluted with H<sub>2</sub>O (117ml) and stirred, under N<sub>2</sub> at 55-60°, while ground *p*-toluenesulfonyl chloride (95.3g, 0.5mole, freed from any acid by dissolving in \*C<sub>6</sub>H<sub>6</sub>, washing with 5% aqueous NaOH, drying with Na<sub>2</sub>SO<sub>4</sub>, evaporating and distilling at 146°/15mm) at such a rate as not to allow the temperature to rise above 60°. The deep yellow colour of the solution disappears after 90g is added and the chloride ceases to dissolve. The mixture is filtered rapidly with suction through a warmed funnel, and the filtrate is set aside at 0-5° for many hours. The solid is filtered off, dissolved in hot 80% EtOH (200ml), filtered to remove traces of S, cooled for many hours at 0-5°, and the white crystals of *potassium p*-toluenethiosulfonate (48-55g, 42-49%) are collected and air dried. [Woodward et al. *Org Synth Coll Vol V 1016 1988, Org Synth 54 33 1974, Beilstein 11 IV 482.*]

**Tolhydroquinone (2-methylbenzene-1,4-dione)** [95-71-6] **M 124.1, m 128-129°, pK<sub>1</sub><sup>20</sup> 10.15, pK<sub>2</sub><sup>20</sup> 11.75**. Crystallise the quinone from EtOH. [Beilstein 6 IV 5866.]

***o*-Toluic acid** [118-90-1] **M 136.2, m 102-103°, b 258-259°/760mm, pK<sup>25</sup> 3.91**. Crystallise the acid from \*benzene (2.5ml/g) and dry in air. The *S*-benzylisothiuronium salt has **m 146°** (from aqueous EtOH). [Beilstein 9 IV 1697.]

***m*-Toluic acid** [99-04-7] **M 136.2, m 111-113°, b 263°/760mm, pK<sup>25</sup> 4.27.** Crystallise the acid from water. [*Beilstein* 9 IV 1712.] Aromatic acid impurities (to <0.05%) can be removed via the (±)- $\alpha$ -methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [*Ley & Yates Organic Process Research & Development* 12 120 2008]. The *S*-benzylisothiuronium salt has **m 140°** (from aqueous EtOH).

***p*-Toluic acid** [99-94-5] **M 136.2, m 178.5-179.5°, b 274-275°/760mm, pK<sup>25</sup> 4.37.** Crystallise the acid from water, water/EtOH (1:1), MeOH/water or \*benzene. [*Beilstein* 9 IV 1724.] Aromatic acid impurities (to <0.05%) can be removed via the (±)- $\alpha$ -methylbenzylamine salt as described for 2,4-dichlorobenzoic acid [*Ley & Yates Organic Process Research & Development* 12 120 2008]. The *S*-benzylisothiuronium salt has **m 164°** (from aqueous EtOH).

***o*-Toluidine (2-methylaniline)** [95-53-4] **M 107.2, f -16.3°, b 80.1°/10mm, 200.3°/760mm, d<sub>4</sub><sup>20</sup> 0.999, n<sub>D</sub><sup>20</sup> 1.57246, n<sub>D</sub><sup>25</sup> 1.56987, pK<sup>25</sup> 4.45.** In general, methods similar to those for purifying aniline can be used, e.g. distillation from zinc dust, at reduced pressure, under nitrogen. Berliner and May [*J Am Chem Soc* 49 1007 1927] purified it via the oxalate. Twice-distilled *o*-toluidine is dissolved in four times its volume of diethyl ether, and the equivalent amount of oxalic acid needed to form the dioxalate is added as its solution in diethyl ether. (If *p*-toluidine is present, its oxalate precipitates and can be removed by filtration.) Evaporation of the ethereal solution gives crystals of *o*-toluidine dioxalate [*Beilstein* 12 III 1494, 12 IV 1817]. These are filtered off, recrystallised five times from water containing a small amount of oxalic acid (to prevent hydrolysis), then treated with dilute aqueous Na<sub>2</sub>CO<sub>3</sub> to liberate the amine which is separated, dried (CaCl<sub>2</sub>) and distilled under reduced pressure. The *benzoyl* derivative has **m 144°** (from EtOH). [*Beilstein* 12 H 772, 12 I 372, 12 II 429, 12 III 1837, 12 IV 1744.]

***m*-Toluidine (3-methylaniline)** [108-44-1] **M 107.2, f -30.4°, b 82.3°/10mm, 203.4°/760mm, d<sub>4</sub><sup>20</sup> 0.989, n<sub>D</sub><sup>20</sup> 1.56811, n<sub>D</sub><sup>25</sup> 1.56570, pK<sup>25</sup> 4.71.** It can be purified as for aniline. Twice-distilled, *m*-toluidine is converted to the hydrochloride using a slight excess of HCl, and the salt is fractionally crystallised from 25% EtOH (five times), and from distilled water (twice), rejecting, in each case, the first material that crystallised out. The amine is regenerated and distilled as for *o*-toluidine. The *benzoyl* derivative has **m 125°** (from EtOH). [Berliner & May *J Am Chem Soc* 49 1007 1927, *Beilstein* 12 H 853, 12 I 397, 12 II 463, 12 III 1949, 12 IV 1813.]

***p*-Toluidine (4-methylaniline)** [106-49-0] **M 107.2, m 44.8°, b 79.6°/10mm, 200.5°/760mm, d<sub>4</sub><sup>20</sup> 0.962, n<sub>D</sub><sup>20</sup> 1.5636, n<sub>D</sub><sup>59.1</sup> 1.5534, pK<sup>25</sup> 5.08.** In general, methods similar to those for purifying aniline can be used. It can be separated from the *o*- and *m*-isomers by fractional crystallisation from its melt. *p*-Toluidine has been crystallised from hot water (charcoal), EtOH, \*benzene, petroleum ether or EtOH/water (1:4), and dried in a vacuum desiccator. It can also be sublimed at 30° under vacuum. For further purification, use has been made of the oxalate, the sulfate and acetylation. The oxalate, formed as described for *o*-toluidine, is filtered, washed and recrystallised three times from hot distilled water. The base is regenerated with aqueous Na<sub>2</sub>CO<sub>3</sub> and recrystallised three times from distilled water. [Berliner & May *J Am Chem Soc* 49 1007 1927.] Alternatively, *p*-toluidine is converted to its acetyl derivative which, after repeated crystallisation from EtOH, is hydrolysed by refluxing (50g) in a mixture of 500ml of water and 115ml of conc H<sub>2</sub>SO<sub>4</sub> until a clear solution is obtained. The amine sulfate is isolated, suspended in water, and NaOH is added. The free base is distilled twice from zinc dust under vacuum. The *p*-toluidine is then recrystallised from petroleum ether and dried in a vacuum desiccator or in a vacuum for 6 hours at 40°. The *benzoyl* derivative has **m 158°** (from EtOH). [Berliner & Berliner *J Am Chem Soc* 76 6179 1954, Moore et al. *J Am Chem Soc* 108 2257 1986, *Beilstein* 12 H 880, 12 I 140, 12 II 482, 12 III 2017, 12 IV 1866.]

***p*-Toluidine hydrochloride** [540-23-8] **M 143.6, m 245.9-246.1°.** Crystallise the salt from MeOH containing a few drops of conc HCl or aqueous EtOH. Dry it under vacuum over paraffin chips. [*Beilstein* 12 II 587, 12 III 2021, 12 IV 1869.]

**2-*p*-Toluidinylnaphthalene-6-sulfonic acid** [7724-15-4] **M 313.9, pK<sub>Est</sub> ~0.** Crystallise the acid twice from 2% aqueous KOH and dry it under high vacuum for 4 hours at room temperature. It also crystallises from H<sub>2</sub>O. It is tested for purity by TLC on silica gel with isopropanol as solvent. The free acid is obtained by acidifying a saturated aqueous solution. [*Beilstein* 14 H 762.]

*o*-Tolunitrile [529-19-1] **M 117.2, b 205.2°, d<sub>4</sub><sup>20</sup> 0.992, n<sub>D</sub><sup>20</sup> 1.5279.** Fractionally distil the nitrile, wash it with conc HCl or 50% H<sub>2</sub>SO<sub>4</sub> at 60° until the smell of isonitrile has gone (this also removes any amines), then wash it with saturated NaHCO<sub>3</sub> and dilute NaCl solutions, then dry it with K<sub>2</sub>CO<sub>3</sub> and redistil it. [Beilstein 9 IV 1703.]

*m*-Tolunitrile [620-22-4] **M 117.2, b 209.5-210°/773mm, d<sub>4</sub><sup>20</sup> 0.986, n<sub>D</sub><sup>20</sup> 1.5250.** Dry the nitrile with MgSO<sub>4</sub>, fractionally distil it, then wash it with aqueous acid to remove possible traces of amines, dry and redistil it. [Beilstein 9 H 477, 9 I 191, 9 II 325, 9 III 2324, 9 IV 1717.]

*p*-Tolunitrile [104-85-8] **M 117.2, m 29.5°, b 104-106°/20mm.** Melt the nitrile, dry it with MgSO<sub>4</sub>, fractionally crystallise it from its melt, then fractionally distil it under reduced pressure in a 6-in spinning band column. [Brown *J Am Chem Soc* 81 3232 1959.] It can also be crystallised from \*benzene/petroleum ether (b 40-60°). [Beilstein 9 H 489, 9 I 194, 9 II 330, 9 III 2348, 9 IV 1738.]

**4-Tolyl-2-benzoic acid (4'-methylbiphenyl-2-carboxylic acid)** [7148-03-0] **M 212.2, m 138-139°, 146-148°, pK<sup>25</sup> 3.64.** Crystallise the acid from toluene or \*C<sub>6</sub>H<sub>6</sub> (m 147-148°). [Beilstein 9 H 677, 9 IV 2523.]

*p*-Tolyl carbinol (4-methylbenzyl alcohol) [589-18-4] **M 122.2, m 59.5-60°, 61°, b 116-118°/20mm, 217°/760mm.** Crystallise the alcohol from petroleum ether (b 80-100°, 1g/ml), Et<sub>2</sub>O, pentane or H<sub>2</sub>O (m 61-62.1°). It can also be distilled in a vacuum. [Beilstein 6 H 498, 6 I 248, 6 II 469, 6 III 1779.]

*p*-Tolyl disulfide [103-19-5] **M 246.4, m 45-46°, 168°/45mm.** Purify it by chromatography on alumina using hexane as eluent, then crystallise it from MeOH, and/or distil it in a vacuum. [Kice & Bowers *J Am Chem Soc* 84 2384 1962, Beilstein 6 H 245, 6 I 212, 6 II 400, 6 III 1432, 6 IV 3206.]

**Tolylene-2,4-diisocyanate (toluene-2,4-diisocyanate)** [584-84-9] **M 174.2, m 19.5-21.5°, 20-22°, 28°, b 126°/11mm, 124-126°/18mm, 250°/760mm.** It is purified by fractionation in a vacuum and should be stored in a dry atmosphere. It is soluble in organic solvents but reacts with H<sub>2</sub>O, alcohols (slowly) and amines, all of which could cause explosive polymerisation. It darkens on exposure to light. It has a sharp pungent odour, is **TOXIC** and is **IRRITATING TO THE EYES**. [Siefken *Justus Liebigs Ann Chem* 562 75, 96, 127 1949, Bayer *Angew Chem* 59 257 1947.] It is a reagent for covalent crosslinking of proteins [Wold *Methods Enzymol* 25 623 1972.] [Beilstein 13 IV 243.]

**Tolylene-2,6-diisocyanate (2-methyl-*m*-phenylenediisocyanate)** [91-08-7] **M 174.2, b 129-133°/18mm, d<sup>25</sup> 1.225.** It is purified by fractional distillation in a vacuum. Store it under N<sub>2</sub> in sealed dark ampoules as it is water and light sensitive. Like the preceding 2,4-isomer, it has a sharp pungent odour, is **TOXIC** and is **IRRITATING TO THE EYES**. [Beilstein 13 IV 259.]

*p*-Tolylsulfonylmethyl isocyanide (tosylmethyl isocyanide, TOSMIC) [36635-61-7] **M 195.2, m 114-115°(dec), 116-117°(dec).** Use an efficient fume cupboard. Purify TOSMIC by dissolving (50g) in CH<sub>2</sub>Cl<sub>2</sub> (150ml) and passing it through a column (40×3cm) containing neutral alumina (100g) in CH<sub>2</sub>Cl<sub>2</sub> and eluting with CH<sub>2</sub>Cl<sub>2</sub>. A nearly colourless solution (700ml) is collected, evaporated *in vacuo* and the residue (42-47g) of TOSMIC (m 113-114° dec) is recrystallised once from MeOH (m 116-117° dec). [Hoogenboom et al. *Org Synth* 57 102 1977, Lensen *Tetrahedron Lett* 2367 1972.] It also crystallises from EtOH (charcoal) [Saito & Itano, *J Chem Soc, Perkin Trans 1* 1 1986].

If the reagent had deteriorated considerably it can be synthesised in two steps. *Firstly N-(p-tolylsulfonylmethyl)formamide* is prepared by adding a mixture of aqueous 34-37% formalin (378g, 350ml, ~4.4moles), H<sub>2</sub>O (750ml), excess of formamide (680g, 600ml, 15.5moles) and formic acid (244g, 200ml, 5.3moles) to sodium *p*-toluenesulfinate (267g, 1.5moles), and stirring at 90° to form a clear solution; and then heating at 90-95° is continued for 2 hours (prolonged heating diminishes the yield). The mixture is cooled (ice-salt bath) with continuous stirring, and set aside at -20° overnight after seeding. The white solid is collected, washed by stirring with ice-H<sub>2</sub>O (3×250ml) and drained well. To removed occluded H<sub>2</sub>O, the solid is dissolved in CH<sub>2</sub>Cl<sub>2</sub>, H<sub>2</sub>O is removed (separating funnel), the organic layer is dried (MgSO<sub>4</sub>), filtered, evaporated, and the residue is dried (over P<sub>2</sub>O<sub>5</sub> at 70°) *in vacuo*, to give crude product (134-150g, 42-47%, m 106-110°) which is used in the following step. Pure *N-(p-tolylsulfonylmethyl)formamide*, **m 108-110°**, can be obtained by

recrystallisation from \*C<sub>6</sub>H<sub>6</sub> or 95% aqueous EtOH.

In the *second step*, with exclusion of moisture, POCl<sub>3</sub> (84g, 50ml, 0.55mole) in 1,2-dimethoxyethane (60ml) is added dropwise to a stirred solution of the crude preceding formamide (107g, 0.50mole) in 1,2-dimethoxyethane (250ml), anhydrous Et<sub>2</sub>O (100ml) and Et<sub>3</sub>N (255g, 350ml, 2.52moles) cooled to -5° (ice-salt bath), at such a rate as to keep the temperature between -5° and 0° (requires ~1 hour). During the reaction, the formamide dissolves, white Et<sub>3</sub>NH<sup>+</sup> salts separate, and at the end of the reaction the suspension turns brown in colour. If the mixture is not brownish in colour then more POCl<sub>3</sub> needs to be added. Finally the mixture is stirred for 30 minutes at 0° and ice-H<sub>2</sub>O (1.5L) is added to give a clear brown solution before a fine brown crystalline solid separates. Stir the mixture for 30 minutes at 0°, collect the solid, wash it with cold H<sub>2</sub>O (250ml), dissolve the solid in warm \*C<sub>6</sub>H<sub>6</sub> (400ml at 40-60°), separate the H<sub>2</sub>O, dry the organic layer (MgSO<sub>4</sub>), filter, add charcoal (2g), heat at 60° for 5 minutes, filter, and add petroleum ether (1L, b 40-60°) with swirling. After 30 minutes, the solid is filtered off, dried *in vacuo* to give crude TOSMIC (74-82g, 76-84%) m 111-114° (dec), which can be used for most purposes. Analytically pure product can be obtained by a *third* purification through neutral Al<sub>2</sub>O<sub>3</sub> as stated above. It has IR (Nujol) with  $\nu_{\max}$  at 2150 (N=C), 1320 and 1155 (SO<sub>2</sub>) cm<sup>-1</sup>, and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  at 7.7 (q, 4H, C<sub>6</sub>H<sub>4</sub>), 4.6 (s, 2H, CH<sub>2</sub>) and 2.5 (s, 3H, CH<sub>3</sub>). [Hoogenboom et al. *Org Synth Coll Vol VI* 987 1988.] It is a versatile reagent, provides a formaldehyde anion or dianion equivalent, is used for the synthesis of various keto compounds and under reducing conditions amino and methylamino compounds have been prepared from the same intermediates, and it has been used for the preparation of several classes of heterocyclic compounds [van Leusen *Lect Heterocycl Chem* 5 S111 1980, see also a supplementary issue of Vol 17 of *J Heterocycl Chem*].

***p*-Tolyl urea** [622-51-5] **M 150.2, m 181°**. Crystallise the urea from H<sub>2</sub>O (m 186°), EtOH/water (1:1) or aqueous AcOH (m 184°). [Beilstein 12 H 941, 12 I 425, 12 II 512, 12 III 2084, 12 IV 1923.]

**Tribenzylamine** [620-40-6] **M 287.4, m 93-94°, 230°/13mm, pK<sub>Est</sub> <0**. Crystallise the amine from absolute EtOH or petroleum ether. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. The *hydrochloride* has m 226-228° (from EtOH) and the *picrate* has m 191° (from H<sub>2</sub>O or aqueous EtOH). [Beilstein 12 IV 2183.]

**2,4,6-Tribromoacetanilide** [607-93-2] **M 451.8, m 232°, 238-240°, 240°**. Crystallise the anilide from EtOH. [Beilstein 12 II 359, 12 III 1478.]

**2,4,6-Tribromoaniline** [147-82-0] **M 329.8, m 120°, 122°, pK<sub>Est</sub> ~-0.5 (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Crystallise the aniline from MeOH. The *benzenesulfonamide* derivative has m 198°. [Beilstein 12 H 663, 12 I 329, 12 II 358, 12 III 1477, 12 IV 1538.]

***sym*-Tribromobenzene (1,3,5-tribromobenzene)** [626-39-1] **M 314.8, m 122°**. Crystallise it from glacial acetic acid/water (4:1), then wash with chilled EtOH and dry in air. [Beilstein 5 H 213, 5 IV 685.]

**2,4,6-Tribromophenol** [118-79-6] **M 330.8, m 94°, pK<sup>25</sup> 6.00**. Crystallise the phenol from EtOH or petroleum ether. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. [Beilstein 6 IV 1067.]

***sym*-Tri-*tert*-butylbenzene** [1460-02-2] **M 246.4, m 73.4-73.9°**. Crystallise it from EtOH. [Beilstein 5 IV 1206.]

**2,4,6-Tri-*tert*-butylphenol** [732-26-3] **M 262.4, m 129-132°, 131-131.2°, b 131°/1mm, 147°/10mm, 278°/760mm, pK<sup>25</sup> 12.19**. Distil the phenol under reduced pressure and/or recrystallise it from *n*-hexane or several times from 95% EtOH until the EtOH solution is colourless [Balasubramanian & Bruice *J Am Chem Soc* 108 5495 1986]. It has also been purified by sublimation [Yuan & Bruice *J Am Chem Soc* 108 1643 1986, Wong et al. *J Am Chem Soc* 109 3428 1987]. Purification has also been achieved by passage through a silica gel column followed by recrystallisation from *n*-hexane [Kajii et al. *J Phys Chem* 91 2791 1987]. [Beilstein 6 III 2094, 6 IV 3539.]

**Trichloroacetanilide** [2563-97-5] **M 238.5, m 93.5-94°, 95°**. Crystallise the anilide from \*benzene or 90% EtOH (m 93.5-95.5°). [Sukornick *Org Synth Coll Vol V* 1074 1973, Beilstein 12 H 224, 12 I 193, 12 II 142, 12

III 464, 12 IV 377.]

**2,3,4-Trichloroaniline** [634-67-3] **M 196.5, m 67.5°, b 292°/774mm, pK<sub>Est</sub> ~1.3.** Crystallise the aniline from ligroin. The *acetanilide* has **m 120-122°** (from EtOH or \*C<sub>6</sub>H<sub>6</sub>). [Beilstein 12 H 626.]

**2,4,5-Trichloroaniline** [636-30-6] **M 196.5, m 96.5°, b 270°/760mm, pK 1.09.** Crystallise the aniline from ligroin. [Beilstein 12 H 627, 12 IV 1277.]

**2,4,6-Trichloroaniline** [634-93-5] **M 196.5, m 78.5°, b 127°/14mm, 262°/746mm, pK<sup>25</sup> 0.03.** Crystallise the aniline from ligroin. The *benzoyl* derivative has **m 174°** (from EtOH). [Beilstein 12 H 627, 12 IV 1281.]

**1,2,3-Trichlorobenzene** [87-61-6] **M 181.5, m 52.6°.** Crystallise it from EtOH. [Beilstein 5 IV 664.]

**1,2,4-Trichlorobenzene** [120-82-1] **M 181.5, m 17°, b 210°.** Separate it from a mixture of isomers by washing with fuming H<sub>2</sub>SO<sub>4</sub>, then water, drying with CaSO<sub>4</sub> and slowly fractionally distilling. [Jensen et al. *J Am Chem Soc* 81 3303 1959, Beilstein 5 IV 664.]

**1,3,5-Trichlorobenzene** [108-70-3] **M 181.5, m 64-65°.** Recrystallise it from dry \*benzene or toluene. [Beilstein 5 IV 666.]

**3,4,5-Trichloro-*o*-cresol (3,4,5-trichloro-2-methylphenol)** [608-92-4] **M 211.5, m 77°, pK<sub>Est</sub> ~7.6.** Crystallise the cresol from petroleum ether.

**2,3,5-Trichloro-*p*-cresol (2,3,5-trichloro-4-methylphenol)** [608-91-3] **M 211.5, m 66-67°, pK<sub>Est</sub> ~6.9.** Crystallise the cresol from petroleum ether. [Datta & Mitter *J Am Chem Soc* 41 2034 1919, Beilstein 6 I 204.]

**2,4,5-Trichloro-1-nitrobenzene (1,2,4-trichloro-5-nitrobenzene)** [89-69-0] **M 226.5, m 57°.** Crystallise it from EtOH. [Beilstein 5 IV 728.]

**3,4,6-Trichloro-2-nitrophenol** [82-62-2] **M 242.4, m 92-93°, pK<sub>Est</sub> ~4.1.** Crystallise the nitro-phenol from petroleum ether or EtOH. [Beilstein 6 III 842.]

**2,4,5-Trichlorophenol** [95-95-4] **M 197.5, m 67°, b 72°/1mm, 248°/740mm, pK<sup>25</sup> 7.0.** Crystallise the phenol from EtOH or petroleum ether. [Beilstein 6 IV 962.]

**2,4,6-Trichlorophenol** [88-06-2] **M 197.5, m 67-68°, b 246°/760mm, pK<sup>25</sup> 6.23.** Crystallise the phenol from \*benzene, EtOH or EtOH/water. [Beilstein 6 IV 1005.]

**3,4,5-Trichlorophenol** [609-19-8] **M 197.5, m 100°, pK<sup>25</sup> 7.84.** Crystallise the phenol from petroleum ether/\*benzene mixture. [Beilstein 6 III 729.]

**2,4,5-Trichlorophenoxyacetic acid (2,4,5-T)** [93-76-5] **M 255.5, m 153°, 155-158°, pK<sup>25</sup> 2.83.** Crystallise this herbicide from \*benzene. [Beilstein 6 III 721.] (CANCER SUSPECT)

**1,2,4-Triethylbenzene** [877-44-1] **M 162.3, b 96.8-97.1°/12.8mm, d<sub>4</sub><sup>20</sup> 0.8738, n<sub>D</sub><sup>20</sup> 1.5015.** For separation from a commercial mixture see Dillingham and Reid [*J Am Chem Soc* 60 2606 1938]. [Beilstein 5 IV 1133.]

**1,3,5-Triethylbenzene** [102-25-0] **M 162.3, b 102-102.5°, d<sub>4</sub><sup>20</sup> 0.8631, n<sub>D</sub><sup>20</sup> 1.4951.** For separation from a commercial mixture see Dillingham and Reid [*J Am Chem Soc* 60 2606 1938]. [Beilstein 5 IV 133.]

**4-(Trifluoromethyl)acetophenone** [709-63-7] **M 188.2, m 31-33°, b 79-81°/9mm.** Purify the ketone by distillation or sublimation *in vacuo*. [Beilstein 7 IV 1404.]

**3-Trifluoromethyl-4-nitrophenol** [88-30-2] **M 162.1, m 81°, b 135-138°/0.01mm, pK<sub>Est</sub> ~6.1.** Crystallise

the nitrophenol from \*benzene or from petroleum ether/\*benzene mixture. [*Beilstein* 6 III 1328.]

**$\alpha,\alpha,\alpha$ -Trifluorotoluene (benzotrifluoride)** [98-08-8] M 144.1, b 102.5°,  $d_4^{20}$  1.190,  $n_D^{30}$  1.4100. Purify benzotrifluoride by repeated treatment with boiling aqueous  $\text{Na}_2\text{CO}_3$  (until no test for chloride ion is obtained), dry it with  $\text{K}_2\text{CO}_3$ , then with  $\text{P}_2\text{O}_5$ , and fractionally distil it. [*Beilstein* 5 IV 802.]

**2,3,4-Trihydroxybenzoic acid** [610-02-6] M 170.1, m 207-208°,  $\text{pK}_{\text{Est}(1)} \sim 3.4$ ,  $\text{pK}_{\text{Est}(2)} \sim 7.8$ ,  $\text{pK}_{\text{Est}(3)} > 12$ . Crystallise the acid from water. [*Beilstein* 10 IV 1971.]

**2,4,6-Trihydroxybenzoic acid** [83-30-7] M 170.1, m 205-212°(dec),  $\text{pK}_{\text{Est}(1)} \sim 1.5$ ,  $\text{pK}_{\text{Est}(2)} \sim 8.0$ ,  $\text{pK}_{\text{Est}(3)} > 12$ . Crystallise the acid from water. [*Beilstein* 10 IV 1987.]

**3,4,5-Triiodobenzoic acid** [2338-20-7] M 499.8, m 289-290°, 293°,  $\text{pK}^{25}$  0.65. Crystallise the acid from aqueous EtOH or water. [*Beilstein* 9 H 367, 9 III 1475.]

**3,4,5-Triiodobenzyl chloride** [52273-54-8] M 504.3, m 138°. Crystallise the chloride from  $\text{CCl}_4$ /petroleum ether (charcoal).

**Trimellitic (benzene-1,2,4-tricarboxylic) acid** [528-44-9] M 210.1, m 218-220°,  $\text{pK}_1^{25}$  2.42,  $\text{pK}_2^{25}$  3.71,  $\text{pK}_3^{25}$  5.01. Crystallise the acid from acetic acid or aqueous EtOH. [*Beilstein* 9 IV 3746.]

**1,2,3-Trimethoxybenzene** [634-36-6] M 168.2, m 45-46°. Sublime it under vacuum. [*Beilstein* 6 H 1081, 6 I 540, 6 II 1066, 6 III 6265, 6 IV 7329.]

**1,3,5-Trimethoxybenzene** [621-23-8] M 168.2, m 53°. Sublime it under vacuum. [*Beilstein* 6 III 635, 6 IV 7362.]

**3,4,5-Trimethoxyphenol (Antiarol)** [642-71-7] M 184.2, m 146°, 145-149°,  $\text{pK}_{\text{Est}} \sim 9.4$ . Recrystallise the phenol from 10 times its weight of  $\text{H}_2\text{O}$  (white needles, m 148°). The *acetyl* derivative crystallises as elongated prisms from EtOH with m 74°. [Chapman et al. *J Chem Soc* 3028 1927, Shriner et al. *J Am Chem Soc* 61 2325 1939, *Beilstein* 6 H 1154, 6 II 1118, 6 III 6656.]

**1,2,4-Trimethylbenzene (pseudocumene)** [95-63-6] M 120.2, m -43.8°, b 51.6°/10mm, 167-168°/760mm,  $d_4^{20}$  0.889,  $n_D^{20}$  1.5048. Reflux pseudocumene over sodium and distil it under reduced pressure. [*Beilstein* 6 H 1088, 6 I 542, 6 II 1072, 6 III 6278, 6 IV 7339.]

**2,4,6-Trimethylbenzoic acid (mesitoic acid)** [480-63-7] M 164.2, m 155°,  $\text{pK}^{25}$  3.45. Crystallise mesitoic acid from water, ligroin or carbon tetrachloride [Ohwada et al. *J Am Chem Soc* 108 3029 1986]. [*Beilstein* 9 H 553, 9 I 214, 9 II 360, 9 III 2489, 9 IV 1854.]

**Trimethyl-1,4-benzoquinone** [935-92-2] M 150.1, m 29-30°, 36°, b 98°/10mm, 108°/18mm. Distil the quinone in a vacuum or sublime it *in vacuo* before use. [Smith et al. *J Am Chem Soc* 60 318 1939, *Beilstein* 7 H 161, 7 III 3407, 7 IV 2098.]

**1,3-S,S'-Trimethylene di(*p*-toluenethiosulfonate) [1,3-di(*p*-tosylthio)propane, 1,3-trimethylene di(thio-tosylate)]** [3866-79-3] M 416.6, m 64-67°. This reagent is prepared from potassium thiosulfonate (40g, 0.18 mole, see above [28519-50-8]) and 1,3-dibromopropane (20g, 0.1 mole) which are added to a solution of 95% EtOH (150ml) containing KI (10-20mg, to activate the dibromide), and the stirred mixture is boiled under reflux in the dark under  $\text{N}_2$  for 8 hours. After cooling to  $\sim 25^\circ$ , the mixture is diluted with an equal volume of cold  $\text{H}_2\text{O}$  and shaken. After the mixture settles, the supernatant is removed by decantation, the pale yellow residual oil is washed with cold  $\text{H}_2\text{O}$  (3  $\times$  200ml), once with cold 95% EtOH (100ml) and once with cold absolute EtOH (100ml) also by decantation. The oil is then dissolved in  $\text{Me}_2\text{CO}$  (10ml), diluted with hot absolute EtOH (80ml), and stirred under  $\text{N}_2$  at 0°. If some oil separates, it should be re-dissolved by adding enough  $\text{Me}_2\text{CO}$  (say  $\sim 5$ ml). The solution is seeded with crystals (a small portion of the oil stored at  $-30^\circ$  for a few days crystallises and is

used as seed\*), stirred under N<sub>2</sub> at 0° for 1 hour then for several hours at -30°. The microcrystals are collected (20.2g) and have **m 63.5-65.0°**. This can be used for reactions; but if purer material is required then it should be recrystallised three times from EtOH (180ml, 9 parts) to provide white needles (17.2g, 41%) with **m 66-67°**. The *thioester* always separates as an oil when solutions are cooled to room temperature.

\*If it is difficult to crystallise the ester then seed crystals could be obtained by chromatography through a column of Woelm neutral alumina (Activity Grade I) and eluting with \*C<sub>6</sub>H<sub>6</sub>. The centre fractions giving solid with **m 67°** are combined and recrystallised from EtOH (9parts) to give the pure *trimethylene dithiotosylate* with **m 67.5°**. On a Durapak-Carbowax 400/Poracil C column (Waters Associates, 3ft x 0.125in) with CHCl<sub>3</sub> as eluent it exhibits a single peak. It has IR (CHCl<sub>3</sub>) with  $\nu_{\max}$  at 810 (m), 1015 (w), 1075 (s), 1140 (s), 1180 (w), 1300 (m), 1325 (s), 1410 (w), 1440 (w), 1490 (w), 1590 (w), 2930 (w), and 3030 (w) cm<sup>-1</sup>; and the <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has  $\delta$  at 1.98 (quintet, *J* = 7Hz, 2H, CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>), 2.43 (s, 6H, 2CH<sub>3</sub>), 2.97 (t, *J* = 7Hz, 4H, CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>), 7.30 (d, *J* = 9Hz, 4H, Arom) and 7.75 (d, *J* = 9Hz, 4H, Arom). [Woodward et al. *Org Synth Coll Vol V* 1016 1988, *Org Synth* **54** 33 1974.]

This reagent is useful for the synthesis of 1,3-dithianes from carbonyl compounds and activated methylene groups. Thus a ketone (0.163mmol), reagent (0.218mmol) and dry NaOK (0.907mmol) in EtOH (6ml) are refluxed for 6 hours, dilute with CH<sub>2</sub>Cl<sub>2</sub>, wash with aqueous NaHCO<sub>3</sub>, evaporate extract and purify the residue through a silica gel column by eluting with 1:2 EtOAc/hexane containing 2% of *N*-methylpyrrolidine to give the desired 1,3-dithiane [McMurray et al. *J Org Chem* **49** 3803 1984, see also Saksena & Ganguli *Tetrahedron Lett* **22** 5227 1981]. It also forms 1,3-dithianes, e.g. 2,2-(trimethylenedithio)cyclohexanone, from the respective enamine, i.e. 1-pyrrolidinecyclohex-1-ene [Woodward et al. *Org Synth* **54** 39 1974].

**Trimethyl-1,4-hydroquinone (2,3,5-trimethylbenzene-1,4-diol)** [700-13-0] **M 152.2, m 173-174°, pK<sub>Est(1)</sub>~ 11.1, pK<sub>Est(2)</sub>~ 12.7**. Recrystallise the hydroquinone from water, under anaerobic conditions. [*Beilstein* **6** H 931, **6** IV 5997.]

**2,3,5-Trimethylphenol** [697-82-5] **M 136.2, m 95-96°, b 233°/760mm, pK<sup>25</sup> 10.67**. Crystallise the phenol from water or petroleum ether. [*Beilstein* **6** IV 3248.]

**2,4,5-Trimethylphenol** [496-78-6] **M 136.2, m 70.5-71.5°, pK<sup>25</sup> 10.57**. Crystallise the phenol from water. [*Beilstein* **6** H 509, **6** I 255, **6** II 482, **6** III 1831, **6** IV 3247.]

**2,4,6-Trimethylphenol** [527-60-6] **M 136.2, m 69°, b 220°/760mm, pK<sup>25</sup> 10.86**. Crystallise the phenol from water and sublime it *in vacuo*. [*Beilstein* **6** IV 3253.]

**3,4,5-Trimethylphenol** [527-54-8] **M 136.2, m 107°, b 248-249°/760mm, pK<sup>25</sup> 10.25**. Crystallise the phenol from petroleum ether. [*Beilstein* **6** IV 3245.]

**Trimethylphenylammonium benzenesulfonate** [16093-66-6] **M 293.3**. Crystallise it repeatedly from MeOH (charcoal). [*Beilstein* **12** IV 249.]

**2,4,6-Trinitroanisole** [606-35-9] **M 243.1, m 68°**. Crystallise it from EtOH or MeOH. Dry it *in vacuo*. [*Beilstein* **6** H 288, **6** I 140, **6** II 280, **6** III 968, **6** IV 1456.]

**1,3,6(1,2,4)-Trinitrobenzene** [99-35-4] **M 213.1, m 122-123°**. Crystallise it from glacial acetic acid, CHCl<sub>3</sub>, CCl<sub>4</sub>, EtOH aqueous EtOH or EtOH/\*benzene, after (optionally) heating with dilute HNO<sub>3</sub>. Dry it in air. Fuse it and crystallise it under vacuum. [*Beilstein* **5** H 271, **5** I 140, **5** II 203, **5** III 643, **5** IV 754.]

**2,4,6-Trinitrobenzenesulfonic acid hydrate (TNBS, picrylsulfonic acid)** [2508-19-2] **M 293.2, m 180°,  $\lambda_{\max}$  240nm ( $\epsilon$  650 M<sup>-1</sup>cm<sup>-1</sup>), pK<sub>Est</sub>~ <0**. It is also available as 0.1M and 5%w/v solutions in H<sub>2</sub>O. Recrystallise TNBS from 1M HCl, or a mixture of EtOH (50ml), H<sub>2</sub>O (30ml) and conc HCl (70ml) for 65g of acid, and dry it at 100°. The *diethanolamine salt* had **m 182-183°** [Golumbic *J Org Chem* **11** 518 1946]. [*Beilstein* **11** III 161.]

**2,4,6-Trinitrobenzoic acid** [129-60-8] **M 225.1, m 227-228°, pK<sup>25</sup> 0.65.** Crystallise the acid from distilled H<sub>2</sub>O. Dry in a vacuum desiccator. The *amide* has **m 264°** (from EtOH). [*Beilstein* 9 H 417, 9 I 168, 9 II 285, 9 III 1956, 9 IV 1362.]

**2,4,6-Trinitro-*m*-cresol** [602-99-3] **M 243.1, m 107.0-107.5°, pK<sup>25</sup> 2.8.** Crystallise the cresol successively from H<sub>2</sub>O, aqueous EtOH and \*benzene/cyclohexane, then dry at 80° for 2 hours. [Davis & Paabo *J Res Nat Bur Stand* 64A 533 1960, *Beilstein* 6 H 387, 6 I 194, 6 II 363, 6 III 1331, 6 IV 2079.]

**2,4,7-Trinitro-9-fluorenone** [129-79-3] **M 315.2, m 176°.** Crystallise it from nitric acid water (3:1), wash it with water and dry it under vacuum over P<sub>2</sub>O<sub>5</sub>, or recrystallise it from dry \*benzene. [*Beilstein* 7 II 410, 7 III 2348, 7 IV 1638.]

**2,4,6-Trinitrotoluene (TNT)** [118-96-7] **M 227.1, m 81.0-81.5°.** Crystallise TNT from \*benzene and EtOH. Then fuse (**CARE**) and allow to crystallise under vacuum. Gey, Dalbey and Van Dolah [*J Am Chem Soc* 78 1803 1956] dissolved TNT in acetone and added cold water (1:2:15), the precipitate was filtered off, washed free from solvent and stirred with five parts of aqueous 8% Na<sub>2</sub>SO<sub>3</sub> at 50-60° for 10 minutes. This was filtered, washed with cold water until the effluent was colourless, and air dried. The product was dissolved in five parts of hot CCl<sub>4</sub>, washed with warm water until the washings were colourless and TNT was recovered by cooling and filtering. It was recrystallised from 95% EtOH and carefully dried over H<sub>2</sub>SO<sub>4</sub>. The dry solid should not be heated without taking precautions for a possible **EXPLOSION**. Work with small quantities. [*Beilstein* 5 H 347, 5 I 172, 5 II 268, 5 III 767, 5 IV 873.]

**2,4,6-Trinitro-*m*-xylene** [632-92-8] **M 241.2, m 182.2°.** Crystallise the xylene from ethyl methyl ketone. [*Beilstein* 5 H 381, 5 I 185, 5 II 295, 5 III 845, 5 IV 950.]

**Triphenylamine** [603-34-9] **M 245.3, m 127.3-127.9°, pK<sup>25</sup> -5.0 (in fluorosulfuric acid).** Crystallise the amine from EtOH or from \*benzene/absolute EtOH, diethyl ether and petroleum ether. It is sublimed under vacuum and carefully dried in a vacuum line. Store it in the dark under nitrogen. [*Beilstein* 12 IV 276.]

**1,3,5-Triphenylbenzene** [612-71-5] **M 306.4, m 171-172°, 173-175°, 176°, d<sub>4</sub><sup>0</sup> 1.205.** Purify it by chromatography on alumina using \*benzene or petroleum ether as eluents. Crystallise the triphenylbenzene from EtOH (**m 174°**). [*Beilstein* 5 H 737, 5 I 370, 5 II 670, 5 III 2563, 5 IV 2732.]

**Triphenylene** [217-59-4] **M 228.3, m 198°, b 425°.** Purify triphenylene by zone refining or crystallisation from EtOH or CHCl<sub>3</sub> and sublime. [*Beilstein* 5 IV 2556.]

**1,2,3-Triphenylguanidine** [101-01-9] **M 287.3, m 144°, pK<sup>25</sup> 9.10.** Crystallise the guanidine from EtOH or EtOH/water. Dry it *in vacuo*. [*Beilstein* 12 H 451, 12 I 261, 12 II 246, 12 III 907, 12 IV 866.]

**Triphenylmethane** [519-73-3] **M 244.3, m 92-93°.** Crystallise triphenylmethane from EtOH or \*benzene (with one molecule of \*benzene of crystallisation which is lost on exposure to air or by heating on a water bath). It can also be sublimed under vacuum. It has been given a preliminary purification by refluxing with tin and glacial acetic acid, then filtered hot through a glass sinter disc, and precipitated by addition of cold water. [*Beilstein* 5 H 698, 5 IV 2495.]

**Triphenylmethanol (triphenylcarbinol)** [76-84-6] **M 260.3, m 164°, b 360-380° (without dec), pK<sup>25</sup> -6.63 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Crystallise the carbinol from EtOH, MeOH, CCl<sub>4</sub> (4ml/g), \*benzene, hexane or petroleum ether (b 60-70°). Dry it at 90°. [Ohwada et al. *J Am Chem Soc* 108 3029 1986, *Beilstein* 6 IV 5014.]

**Triphenylmethyl chloride (trityl chloride)** [76-83-5] **M 278.9, m 111-112°, b 230-235°/20mm.** Crystallise trityl chloride from iso-octane. Also crystallise it from 5 parts of petroleum ether (b 90-100°) and 1 part of acetyl chloride using 1.8g of solvent per g of chloride. Dry it in a desiccator over soda lime and paraffin wax. [Bachman *Org Synth Coll Vol III* 841 1955, Thomas & Rochow *J Am Chem Soc* 79 1843 1957, Moisel et al. *J Am Chem Soc* 108 4706 1986, *Beilstein* 5 H 750, 5 I 346, 5 II 615, 5 III 2315, 5 IV 2497.] It is moisture

sensitive, store in a well sealed container. **LACHRYMATORY.**

***dl*-Tropic [(3-hydroxy-2-phenylpropionic) acid, Tropaic acid]** [529-64-6] **M 166.2, m 116-117°, 118°, pK<sup>25</sup> 4.12.** Tropic acid crystallises from H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub> in needles or plates. It is the (±)-acid component (*dl*-troyl ester) of atropine [51-55-8]. It is best prepared from acetophenone cyanohydrin *via* atrolactic acid [515-30-0], atropic acid, *chlorohydratropic acid* followed by hydrolysis of the latter in boiling aqueous 1.73N Na<sub>2</sub>CO<sub>3</sub> in high yielding steps. No racemisation occurred when chiral intermediates were used (see next entry). [McKenzie & Wood *J Chem Soc* **115** 828 1919, McKenzie & Strathern *J Chem Soc* **127** 86 1925.] [*Beilstein* **10** IV 664.]

***S*(-)-Tropic acid [(-)3-hydroxy-2-phenylpropionic acid]** [16202-15-6] **M 166.2, m 128-129°, [α]<sub>D</sub><sup>13</sup> -72.5° (c 2.6, EtOH), [α]<sub>D</sub><sup>15</sup> -79.0° (c 1.5, H<sub>2</sub>O), [α]<sub>D</sub><sup>15</sup> -83.3° (c 1.8, Me<sub>2</sub>CO).** It is the (-)-acid component of *l*-hyoscyamine (i.e. *l*-troyl atropine, [101-31-5]). It has been prepared by optical resolution of (±)-tropic acid using the quinine diastereomeric salt for obtaining (+)-tropic acid, and then with the morphine diastereomeric salt for obtaining the natural (-)-tropic acid. The salts were hydrolysed with dilute H<sub>2</sub>SO<sub>4</sub>, extracted thoroughly into Et<sub>2</sub>O, dried (Na<sub>2</sub>SO<sub>4</sub>), evaporated, and the residue was recrystallised from \*C<sub>6</sub>H<sub>6</sub> (in which it is sparingly soluble) to give lustrous needles m 128-129°. It also crystallises from H<sub>2</sub>O in glassy needles and plates grouped in rosettes. The enantiomer, obtained in the same way, has identical properties except for opposite rotations. [McKenzie & Wood *J Chem Soc* **115** 828 1919, McKenzie & Strathern *J Chem Soc* **127** 86 1925]. The absolute configuration of the (-)-tropic acid was deduced as *S*- by conversion into (+)-alanine. [Fedor & Csepregy *J Chem Soc* 3222 1961, for enzymic kinetic resolution see Klomp et al. *Tetrahedron Asymmetry* **16** 3892 2005.]

**Tropolone (2-hydroxycyclohepta-2,4,6-trien-1-one)** [533-75-5] **M 122.1, m 49-50°, b 81-84°/0.1mm, pK<sub>1</sub> -0.53 (protonation of CO, aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub> 6.67 (acidic OH).** Crystallise tropolone from hexane or petroleum ether and sublime it at 40°/4mm. Also distil it at high vacuum. [*Beilstein* **8** IV 159.]

**Tyramine (4-hydroxybenzylamine)** [51-67-2] **M 137.2, m 164-165°, b 171-181°/8mm, pK<sub>1</sub><sup>25</sup> 9.74 (OH), pK<sub>2</sub><sup>25</sup> 10.52 (NH<sub>2</sub>).** Crystallise tyramine from \*benzene or EtOH. [*Beilstein* **13** IV 1788.]

**Tyramine hydrochloride** [60-19-5] **M 173.6, m 274-276°.** Crystallise the hydrochloride from EtOH by addition of diethyl ether, or from conc HCl. [*Beilstein* **3** II 355.]

**Vanillin (4-hydroxy-3-methoxybenzaldehyde)** [121-33-5] **M 152.2, m 83°, b 170°/15mm, pK<sup>25</sup> 7.40.** Crystallise vanillin from water or aqueous EtOH, or by distillation *in vacuo*. [*Beilstein* **8** IV 1763.]

**Veratraldehyde (vanillin methyl ether)** [120-14-9] **M 166.2, m 42-43°.** Crystallise the ether from diethyl ether, petroleum ether, CCl<sub>4</sub> or toluene. [*Beilstein* **8** IV 1765.]

**Variamine Blue RT [4-(phenylamino)benzenediazonium sulfate (1:1)]** [4477-28-5] **M 293.3, m 120°(dec), CI 37240, λ<sub>max</sub> 377 nm.** Dissolve 10g of the dye in 100ml of hot water. Sodium dithionite (0.4g) is then added, followed by active carbon (1.5g) and is filtered hot. To the colourless or slightly yellow filtrate a solution of saturated NaCl is added, and the mixture is cooled. The needles are filtered off, washed with cold water, dried at room temperature, and stored in a dark bottle (light sensitive). [Anderson & Steedly *J Am Chem Soc* **76** 5144 1954, Erdey *Chem Analyst* **48** 106 1959, *Beilstein* **16** H 602, **16** I 371, **16** II 307, **16** III 575.]

**9-Vinylanthracene** [2444-68-0] **M 204.3, m 65-67°, b 61-66°/10mm.** Purify it by vacuum sublimation. It has also been purified by chromatography on silica gel with cyclohexane as eluent, and recrystallised from EtOH [Werst et al. *J Am Chem Soc* **109** 32 1987]. [*Beilstein* **5** IV 2415.]

**4-Vinylbenzyl chloride (4-chloromethylstyrene)** [1592-20-7] **M 152.6, b 58-62°/0.4mm, 75.5-79°/2mm, 229°/760mm, d<sup>25</sup> 1.083, n<sub>D</sub><sup>20</sup> 1.572.** Purify 4-vinylbenzyl chloride by dissolving it in Et<sub>2</sub>O, washing it with 0.5% of aqueous NaOH, separating, drying the organic layer (Na<sub>2</sub>SO<sub>4</sub>), evaporating and distilling the residual oil under N<sub>2</sub> *in vacuo*. Add 0.05% of 4-*tert*-butylcatechol as stabiliser. It is **lachrymatory**. [Nishikubo et al. *Tetrahedron Lett* **22** 3872 1981, Tanimoto et al. *Synth Commun* **4** 193 1974, *Beilstein* **6** IV 3818.]

**1-Vinylnaphthalene** [826-74-4] **M 154.2, b 68-69°/0.4mm, 100-101°/3.7mm, 124-125°/15mm, d 1.040, n<sub>D</sub><sup>20</sup> 1.653.** Commercial material can contain up to 2% of poly-1-vinylnaphthalene. Fractionally distil it under reduced pressure using a spinning-band column, dry it with CaH<sub>2</sub> and again distil it under vacuum. Store it in sealed ampoules in a freezer. It has λ<sub>max</sub> (cyclohexane) at 210 and 320nm. The *picrate* has **m 105-106°**. [Beilstein 5 H 585, 5 III 1773, 5 IV 1833.] It is an **irritant**.

**2-Vinylnaphthalene** [827-54-3] **M 154.2, m 66°, 65-68°, b 76-81°/2.5mm, 135-137°/18mm.** Commercial material can contain up to 5% of MeOH. The flammable solid can be fractionally distilled under reduced pressure using a spinning-band column; dry it with CaH<sub>2</sub> and again distil it under vacuum. It crystallises from aqueous MeOH, EtOH (**m 66°**) or petroleum ether (**m 66-66.5°**). Store it in sealed ampoules in a freezer. The *picrate* has **m 91-92°** (from EtOH). [Beilstein 5 III 1775, 5 IV 1833.] It is an **irritant**.

**Violanthrene (dibenzanthrene, 5,10-dihydroviolanthrene A)** [81-31-2] **M 428.5.** Purify violanthrene by vacuum sublimation over Cu in a muffle furnace at 450°/25mm in a CO<sub>2</sub> atmosphere [Scholl & Meyer *Chem Ber* 67 1229 1934]. [Beilstein 5 I 392.] *Violanthrene A* (anthro[9,1,2-*cde*]benzo[*rst*]pentaphene [188-87-4] **M 426.5** has **m 506°**. [Clar *Chem Ber* 76 458 1943, Beilstein 5 III 2778; *Violanthrene B* (*violanthrone*) Beilstein 7 I 466, 7 II 818, 7 III 4539.]

**Xylene** [1330-20-7] **M 106.1 (mixed isomers).** Usual impurities are ethylbenzene, paraffins, traces of sulfur compounds and water. With a very efficient still, *o*-xylene can be fractionally distilled from a mixture of isomers. Purify (and dry) by fractional distillation from LiAlH<sub>4</sub>, P<sub>2</sub>O<sub>5</sub>, CaH<sub>2</sub> or sodium. This treatment can be preceded by shaking successively with conc H<sub>2</sub>SO<sub>4</sub>, water, aqueous 10% NaOH, water and mercury, and drying with CaCl<sub>2</sub> for several days. [Beilstein 5 H 360.]

***o*-Xylene** [95-47-6] **M 106.2, f -25.2°, b 84°/14mm, 144.4°/760mm, d<sub>4</sub><sup>20</sup> 0.88020, d<sub>4</sub><sup>25</sup> 0.87596, n<sub>D</sub><sup>20</sup> 1.50543, n<sub>D</sub><sup>25</sup> 1.50292.** The general purification methods listed under xylene are applicable [Clarke & Taylor *J Am Chem Soc* 45 831 1923]. *o*-Xylene (4.4Kg) is sulfonated by stirring for 4 hours with 2.5L of conc H<sub>2</sub>SO<sub>4</sub> at 95°. After cooling, and separating the unsulfonated material, the product is diluted with 3L of water and neutralised with 40% NaOH. On cooling, sodium *o*-xylene sulfonate separates and is recrystallised from half its weight of water. [A further crop of crystals is obtained by concentrating the mother liquor to one-third of its volume.] The salt is dissolved in the minimum amount of cold water, then mixed with the same amount of cold water, and with the same volume of conc H<sub>2</sub>SO<sub>4</sub> and heated to 110°. *o*-Xylene is regenerated and steam distils. The distillate is saturated with NaCl, the organic layer is separated, dried and redistilled. [Beilstein 5 H 362, 5 I 179, 5 II 281, 5 III 807, 5 IV 917.]

***m*-Xylene** [108-38-3] **M 106, f -47.9°, b 139.1°/760mm, d<sub>4</sub><sup>20</sup> 0.86417, d<sub>4</sub><sup>25</sup> 0.85990, n<sub>D</sub><sup>20</sup> 1.49721, n<sub>D</sub><sup>25</sup> 1.49464.** The general purification methods listed under *xylene* are applicable. The *o*- and *p*-isomers can be removed by their selective oxidation when a *m*-xylene sample containing them is boiled with dilute HNO<sub>3</sub> (one part conc acid to three parts water). After washing with water and alkali, the product can be steam distilled, collected as for *o*-xylene, then distilled and purified further by sulfonation. [Clarke & Taylor *J Am Chem Soc* 45 831 1923.] *m*-Xylene is selectively sulfonated when a mixture of xylenes is refluxed with the theoretical amount of 50-70% H<sub>2</sub>SO<sub>4</sub> at 85-95° under reduced pressure. By using a still resembling a Dean and Stark apparatus, water in the condensate can be progressively withdrawn while the xylene is returned to the reaction vessel. After cooling, then adding water, unreacted xylenes are distilled off under reduced pressure. The *m*-xylene sulfonic acid is subsequently hydrolysed by steam distillation up to 140°, the free *m*-xylene is washed, dried with silica gel and again distilled. It is stored over molecular sieves Linde type 4A. [Beilstein 5 H 370, 5 I 182, 5 II 287, 5 III 823, 5 IV 932.]

***p*-Xylene** [106-42-3] **M 106.2, f 13.3, b 138.3°/760mm, d<sub>4</sub><sup>20</sup> 0.86105, d<sub>4</sub><sup>25</sup> 0.85669, n<sub>D</sub><sup>20</sup> 1.49581, n<sub>D</sub><sup>25</sup> 1.49325.** The general purification methods listed for *xylene* above are applicable. *p*-Xylene can readily be separated from its isomers by crystallisation from such solvents as MeOH, EtOH, isopropanol, acetone, butanone, toluene, pentane or pentene. It can be further purified by fractional crystallisation by partial freezing, and stored over sodium wire or molecular sieves Linde type 4A. [Stokes & French *J Chem Soc, Faraday Trans 1* 76 537 1980, Beilstein 5 H 382, 5 I 185, 5 II 296, 5 III 845, 5 IV 951.]

## HETEROCYCLIC COMPOUNDS

**Acetaldehyde ammonia trimer (hexahydro-2,4,6-trimethyl-1,3,5-triazine trihydrate)** [76231-37-3] **M 183.3, m 94-96°, 95-97°, 97°, b 110°(partly dec)**. It crystallises from EtOH/Et<sub>2</sub>O. When prepared, it separates as the *trihydrate* which can be dried in a vacuum over CaCl<sub>2</sub> at room temperature to give the *anhydrous* compound with the same melting point. The *dihydrate* melts at 25-28°, then resolidifies and melts again at 94-95°. *IT IRRITATES THE EYES AND MUCOUS MEMBRANES*. [Nielson et al. *J Org Chem* **38** 3288 1973, *Beilstein* **26** III/IV 25.]

**2-Acetamido-5-nitrothiazole (Acinitrazole)** [140-40-9] **M 187.2, m 264-265°**. Recrystallise acinitrazole from EtOH or AcOH. [Hurd & Wehrmeister *J Am Chem Soc* **71** 4007 1949, *Beilstein* **27** III/IV 4676.]

**4-Acetamido-2,2,6,6-tetramethylpiperidine-1-oxyl (acetamidoTEMPO)** [14691-89-5] **M 213.3, m 144-146°, 146-147°**. Dissolve the 1-oxyl in CH<sub>2</sub>Cl<sub>2</sub>, wash it with saturated K<sub>2</sub>CO<sub>3</sub>, then saturated aqueous NaCl, dry (Na<sub>2</sub>SO<sub>4</sub>), filter and evaporate. The red solid is recrystallised from aqueous MeOH, **m 147.5°**. [Ma & Bobbitt *J Org Chem* **56** 6110 1991, Rozantsev & Kokhanov *Bull Acad Sci USSR, Div Chem Sci* **15** 1422 1966, *Beilstein* **22/8** V 174.]

**5-Acetamido-1,3,4-thiadiazole-2-sulfonamide (Acetazolamide)** [59-66-5] **M 222.3, m 256-259° (dec)**. It is recrystallised from water. [Roblin & Clapp *J Am Chem Soc* **72** 4890 1950, *Beilstein* **27** III/IV 8219.]

**Acetoacetylpipepidide** [1128-87-6] **M 169.2, b 88.9°/0.1mm, n<sub>D</sub><sup>20</sup> 1.4983**. Dissolve it in \*benzene, extract with 0.5M HCl to remove basic impurities, wash with water, dry, and distil it at 0.1mm [Wilson *J Org Chem* **28** 314 1963]. [*Beilstein* **20** IV 1121.]

**4-Acetoxy-2-azetidinone** [28562-53-0] **M 129.1, m 38-41°**. Dissolve it in CHCl<sub>3</sub>, dry (MgSO<sub>4</sub>), concentrate at 40°/70mm, or better at room temperature to avoid decomposition. Wash and stir the residual oil with hexane by decantation and discard the wash. Dry the oil at high vacuum when it should solidify, **m 34°**. It can be distilled at high vacuum, 80-82°/10<sup>-3</sup>mm, but this results in extensive losses. The purity can be checked by TLC using Merck Silica Gel F<sub>254</sub> and eluting with EtOAc. The azetidinone has R<sub>F</sub> 0.38 (typical impurities have R<sub>F</sub> 0.67). The spots can be detected by the **TDM spray**. This is prepared from (A) 2.5g of 4,4'-tetramethyldiaminodiphenylmethane (TDM) in 10ml AcOH and diluted with 50ml of H<sub>2</sub>O, (B) 5g KI in 100ml of H<sub>2</sub>O and (C) 0.3g ninhydrin in 10ml of AcOH and 90ml of H<sub>2</sub>O. The spray is prepared by mixing (A) and (B) with 1.5ml of (C) and stored in a brown bottle. [Clauß et al. *Justus Liebigs Ann Chem* 539 1974, Michel et al. *Org Synth* **65** 135 1987, *Beilstein* **21/12** V 4.]

**3R,4R,1'R-4-Acetoxy-3-[1-(tert-butylmethylsilyloxy)ethyl]-2-azetidinone**. See in "Miscellaneous As, B, P, Si, S, Se and Te Compounds" in this Chapter.

**N-Acetylcaprolactam (1-acetylazepan-2-one, 1-acetylhexahydro-2H-azepin-2-one)** [1888-91-1] **M 155.2, b 85-86°/0.05mm, 119.5-222°/11mm, 134-135°/26mm, d<sub>4</sub><sup>25</sup> 1.094, n<sub>D</sub><sup>20</sup> 1.489**. Distil the lactam in a vacuum, but if the IR has broad OH bands (*ca*  $\nu_{\max}$  at ~3300cm<sup>-1</sup>), then add *ca* one half volume or more of Ac<sub>2</sub>O, reflux for 2 hours and distil or dissolve it in toluene add MeCOCl heat at 100°/2 hours and distil in a vacuum. Its IR has  $\nu_{\max}$  at 1730cm<sup>-1</sup> (film). [Stoll & Griehl *Helv Chim Acta* **48** 1805 1965, Bose et al. *Tetrahedron* **30** 3 1974, *Beilstein* **21** III/IV 3210, **21/6** V 460.]

**3-Acetyl-2,5-dimethylthiophene** [2530-10-1] **M 154.2, b 62°/0.25mm, 105-108°/15mm, 224°/atm, d<sub>4</sub><sup>25</sup> 1.086, n<sub>D</sub><sup>20</sup> 1.544**. Distil the thiophene in a vacuum, or dissolve it in \*C<sub>6</sub>H<sub>6</sub>, shake this with 6M HCl, then wash with H<sub>2</sub>O, dry the organic layer, evaporate, and distil the residue preferably in a vacuum. [Glaze et al. *J Chem Soc Perkin Trans 1* 957 1985, *Beilstein* **17** H 298.]

**N-Acetylimidazole** [2466-76-4] **M 110.1, m 101.5-102.5°, pK<sup>25</sup> 3.6.** It is recrystallised from isopropenyl acetate and dried in a vacuum over P<sub>2</sub>O<sub>5</sub>. [Riordan & Valee *Methods Enzymol* **25** 500 1972, *Beilstein* **23/4** V 218.]

**1-Acetylimidazole** [576-15-8] **M 159.2, b 123-125°/8mm, 148-150°/10mm, 152-153°/15mm, d<sub>4</sub><sup>25</sup> 1.387, n<sub>D</sub><sup>20</sup> 1.6100.** It has been prepared by phase transfer catalysis of a mixture of AcCl/NaOH and Indole [Illi *Synthesis* 387 1979]. Purify it by distillation in a vacuum, but if it is very discoloured, then distil it in steam, separate the layers, dry the organic layer (Na<sub>2</sub>SO<sub>4</sub>), filter it, and distil it in a vacuum. The *picrate* has **m 98-99°**. *1,3-Diacetylimidazole* is formed when indole is treated with Ac<sub>2</sub>O at 180-200°, and distils with 1-acetylimidazole. The latter distils in steam, whereas the diacetylimidazole remains behind. It is hydrolysed by boiling alkali to yield indole, and can be acetylated further to *1,3-diacetylimidazole* **m 151°**. [For NMR see Elguero et al. *Org Mag Res* **7** 445 1975, Ciamecian & Chem Ber **22** 1977 1889, Zatti & Ferratini *Chem Ber* **23** 1357 1890, Plieninger & Werst *Chem Ber* **89** 2783, 2788 1956, *Beilstein* **20** H 309, **20** II 200, **20** III/IV 3182, **20/7** V 19.]

**3-Acetylimidazole** [703-80-0] **M 159.2, m 189°, 188-190°, 188-192°, 191-193°, 194°, pK<sup>25</sup> 12.99 (acidic NH).** It is formed by reaction of indole-MgCl and AcCl, and also by hydrolysis of 1,3-diacetylimidazole (see above). It is purified by recrystallisation from MeOH, \*C<sub>6</sub>H<sub>6</sub> containing a little EtOH or \*C<sub>6</sub>H<sub>6</sub>; and by sublimation *in vacuo*. The *phenylureido* derivative has **m 154°**. The (*E*)-*oxime* has **m 144-145°** (also 149° reported) (from H<sub>2</sub>O) [Kostyuchenko et al. *J Org Chem USSR (Engl Trans)* **8** 2471 1972], and the (*Z*)-*oxime* has **m 95°** [Kostyuchenko et al. *J Org Chem USSR (Engl Trans)* **8** 2469, 2473 1972]. The (*E*)-*O-methyloxime* crystallises from EtOH with **127-129°**. [Oddo & Sessa *Gazz Chim Ital* **41** 234 1911, Baker *J Chem Soc* 461 1946, *Beilstein* **21** H 316, **21** I 300, **21** II 264, **21** III/IV 3775, **21/8** V 297.]

**4-Acetylmorpholine** [1696-20-4] **M 129.2, m 13.8-14°, 14°, 14.5°, b 96-97°/6mm, 113-128°/22mm, 242-247°/760mm, d<sub>4</sub><sup>20</sup> 1.0963, n<sub>D</sub><sup>20</sup> 1.4830.** Distil it through an 8inch Fenske (glass helices packing) column with a manual take-off head. Purify it by fractional distillation. The *hydrobromide* has **m 172-175°**. [Brace *J Am Chem Soc* **75** 357 1953, deBenneville et al. *J Org Chem* **21** 1072 1956, *Beilstein* **27** III/IV 274.]

**N-Acetyl-D-penicillamine (N-acetyl-3-mercapto-D-valine)** see in "Miscellaneous Compounds", Chapter 7.

**1-Acetylpiperazine** [13889-98-0] **M 128.2, m 32-34°, 52°, pK<sup>25</sup> 7.94.** Purify 1-acetylpiperazine by recrystallisation from 40% aqueous EtOH or from EtOH/Et<sub>2</sub>O. It is an **irritant**, and is *hygroscopic*. The *hydrochloride* has **m 191°** (from EtOH), and the *tosylate salt* has **m 148-149°** (from EtOH/EtOAc, 1:16). The free base, however, cannot be isolated by basifying the tosylate salt and extraction with CH<sub>2</sub>Cl<sub>2</sub>. [Jacobi *Chem Ber* **66** 113 1933, Mosher et al. *J Am Chem Soc* **75** 4949 1953, Hall *J Am Chem Soc* **78** 2570 1956, *Beilstein* **23** IV 1786.]

**1-Acetyl-4-piperidone** [32161-06-1] **M 141.2, b 124-128°/0.2mm, 218°/760mm, d<sub>4</sub><sup>25</sup> 1.1444, n<sub>D</sub><sup>20</sup> 1.5023.** Purify it by fractional distillation through a short Vigreux column (15mm). The *2,4-dinitrophenylhydrazone* has **m 212-213°** (from EtOH). It is freely soluble in H<sub>2</sub>O but insoluble in Et<sub>2</sub>O. [McElvain & McMahon *J Am Chem Soc* **71** 901 1949, *Beilstein* **21/6** V 246.]

**3-Acetylpyridine** [350-03-8] **M 121.1, m 13-14°, b 65-66°/1mm, 92-95°/8-9mm, 105°(113°)/16mm, 219-221°/760mm, d<sub>4</sub><sup>20</sup> 1.1065, n<sub>D</sub><sup>20</sup> 1.1065, pK<sup>25</sup> 3.18.** It is purified by dissolving in HCl, extracting with Et<sub>2</sub>O to remove the possible impurity of nicotinic acid, basified with NaOH and extracted with Et<sub>2</sub>O. The dried extract is filtered, evaporated and the residual oil is distilled. If the NMR spectrum indicates further impurities, then convert it to the *phenylhydrazone* (**m 137°**, yellow needles from EtOH). This is hydrolysed with HCl [Engler & Kiby *Chem Ber* **22** 597 1889], the phenylhydrazine HCl is removed by filtration, NaNO<sub>2</sub> is added, the solution is basified with aqueous NaOH and extracted with Et<sub>2</sub>O as before and distilled at atmospheric pressure to give 3-acetylpyridine as a colourless oil. Purification can also be achieved by shaking with 50% aqueous KOH, extracting with Et<sub>2</sub>O, drying the extract and distilling it at atmospheric pressure or *in vacuo*. [Kloetzel & Chubb *J Am Chem Soc* **79** 4226 1957.] The *hydrochloride* has **m 180-181°** (from MeOH/EtOH), the *picrate* has **m 133.8-134.8°** (from H<sub>2</sub>O), and the *phenylhydrazone* has **m 137°** (129-130°)(needles, from EtOH) [Webb & Webb *J Am Chem Soc* **71** 2285 1949]. The *ketoxime* has **m 112°** (from EtOH or \*C<sub>6</sub>H<sub>6</sub>). [Strong & McElvain

*J Am Chem Soc* **55** 816 1933, Kolloff & Hunter *J Am Chem Soc* **63** 490 1941, *Beilstein* **21/7** V 394.]

**2-Acetylthiazole** [24295-03-2] **M 127.2, b 89-91° (90-95°)/12mm, 95-105°/15mm,  $d_4^{20}$  1.23,  $n_D^{20}$  1.55.** Check NMR spectrum; if it is not too bad, distil it through an efficient column in a vacuum. Otherwise purify it *via* the oxime. The *oxime* sublimes at 140-145°, **m 159°**, and when crystallised from H<sub>2</sub>O has **m 163-165.5°**. [Erlenmeyer et al. *Helv Chim Acta* **31** 1142 1948, Waisvisz et al. *J Am Chem Soc* **79** 4524 1957, Menasseé et al. *Helv Chim Acta* **40** 554 1957, *Beilstein* **27** IV 2617.]

**2-Acetylthiophene (methyl 2-thienyl ketone)** [88-15-3] **M 126.2, m 9.2-10.5°, 10.45°, 10-11°, b 77°/4mm, 89-91°/9mm, 94.5-96.5°/13mm, 213-214°/atm,  $d_4^{20}$  1.17,  $n_D^{20}$  1.5666.** Fractionally distil the thiophene through a 12-plate column, and fraction **b 77°/4mm** is collected. Also wet the acetylthiophene in order to remove and free thiophene which forms an azeotrope with H<sub>2</sub>O, **b 68°**. Store it in a brown bottle, and the clear colourless liquid remains thus for extended periods. [Kosak & Hartough *Org Synth Coll Vol III* 14 1955, Hartough & Kosak *J Am Chem Soc* **69** 3093 1947.] The red *4-nitrophenylhydrazone* crystallises from EtOH with **m 181-182°**. [*Beilstein* **17/9** V 387.]

**3-Acetylthiophene (methyl 3-thienyl ketone, acetothienone)** [1468-83-3] **M 126.2, m 57°, 60-63°, b 106-107°/25mm, 208-210°/748mm.** Recrystallise the thiophene from petroleum ether (b 30-60°) or EtOH. The *2,4-dinitrophenylhydrazone* crystallises from CHCl<sub>3</sub>, **m 265°**, and the *semicarbazone* crystallises from EtOH, **m 174-175°**. [Campaigne & Le Suer *J Am Chem Soc* **70** 1555 1948, *Beilstein* **17/9** V 399.]

**Aconitine** [302-27-2] **M 645.8, m 204°,  $[\alpha]_{546} +20°$  (c 1, CHCl<sub>3</sub>),  $pK^{15}$  8.35.** Crystallise it from EtOH, CHCl<sub>3</sub> or toluene. [*Beilstein* **21/6** V 310.]

**Aconitine hydrobromide** [6034-57-7] **M 726.7, m 207°.** Crystallise the salt from water or EtOH/ether. [*Beilstein* **21/6** V 310.]

**Acridine (2,3-benzoquinoline)** [260-94-6] **M 179.2, m 111° (sublimes), b 346°,  $pK^{25}$  5.58 ( $pK^{25}$  of excited state 10.65).** Acridine has been crystallised twice from \*benzene/cyclohexane, or from aqueous EtOH, then sublimed, removing and discarding the first 25% of the sublimate. The remainder is again crystallised and sublimed, discarding the first 10-15% [Wolf & Anderson *J Am Chem Soc* **77** 1608 1955].

Acridine can also be purified by crystallisation from *n*-heptane and then from ethanol/water after pre-treatment with activated charcoal, or by chromatography on alumina with petroleum ether in a darkened room. *Alternatively*, acridine can be precipitated as the hydrochloride from \*benzene solution by adding HCl, after which the base is regenerated, dried at 110°/50mm, and recrystallised to constant melting point from petroleum ether [Cumper et al. *J Chem Soc* 4518 1962]. The regenerated free base may be recrystallised, chromatographed on basic alumina, then vacuum-sublimed and zone-refined. [Williams & Clarke, *J Chem Soc, Faraday Trans 1* **73** 514 1977, Albert, *The Acridines* Arnold Press 1966.] It can exist in five crystalline forms and is steam volatile. It is a strong IRRITANT to skin and mucous membranes and can become a chronic irritant—handle it with CARE. [*Beilstein* **20/8** V 199.]

**Acridine Orange (3,6-bisdimethylaminoacridine)** [494-38-2] **M 349.94, m 165°(dec), 181-182°, 184-186° (free base),  $pK_1^{20}$  -3.3,  $pK_2^{20}$  0.2,  $pK_3^{20}$  10.1.** The double salt with ZnCl<sub>2</sub> (6g) is dissolved in water (200ml) and stirred with four successive portions (12g each) of Dowex-50 ion-exchange resin (K<sup>+</sup> form) to remove the zinc. The solution is then concentrated in vacuum to 20ml, and 100ml of ethanol is added to precipitate KCl which is removed. Ether (160ml) is added to the solution from which, on chilling, the dye crystallises as its chloride. It is separated by centrifugation, washed with chilled ethanol and ether, and dried under vacuum, before being recrystallised from ethanol (100ml) by adding ether (50ml), and chilling. Yield 1g. [Pal & Schubert *J Am Chem Soc* **84** 4384 1962].

It was recrystallised twice as the free base from ethanol or methanol/water by dropwise addition of NaOH (less than 0.1M). The precipitate was washed with water and dried under vacuum. It was dissolved in CHCl<sub>3</sub> and chromatographed on alumina: the main sharp band was collected, concentrated and cooled to -20°. The precipitate was filtered off, dried in air, then dried for 2 hours under vacuum at 70°. [Stone & Bradley *J Am*

*Chem Soc* **83** 3627 1961, Blauer & Linschitz *J Phys Chem* **66** 453 1962, Albert *J Chem Soc* 244 1947, *Beilstein* **22** III/IV 5490, **22/11** V 326.]

**Acridine Yellow G (2,7-dimethylacridine-3,6-diamine hydrochloride)** [135-49-9] **M 273.8, m 325° (free base), CI 46025, pK<sub>1</sub><sup>20</sup> -3.0, pK<sub>2</sub><sup>20</sup> 0.5, pK<sub>3</sub><sup>20</sup> 8.9.** Note that the hydrochloride is normally called Acridine Yellow G and not the free base. The free base recrystallises from 1:1 \*benzene/methanol, or from 800 parts of EtOH. It is converted to the sparingly soluble hydrochloride with dilute HCl, filtered off and dried. [Albert *J Chem Soc* 248 1947, *Beilstein* **22** III/IV 5499, **22/11** V 340.]

**N-(9-Acridinyl)maleimide (NAM)** [49759-20-8] **M 274.3, m 248°, 255-258°.** Purify NAM by chromatography on silica gel using CH<sub>2</sub>Cl<sub>2</sub> as eluant. Evaporation of pooled fractions that gave the correct NMR spectra gave a solid which was crystallised from Me<sub>2</sub>CO as pale yellow prisms. Its IR(ν<sub>max</sub>) has ν<sub>max</sub> at 1710 (imide); the UV (MeOH) has λ<sub>max</sub> (nm) (ε M<sup>-1</sup>cm<sup>-1</sup>) at 251 (159 500), 343 shoulder (7,700), 360 (12,400) and 382 shoulder (47,000). [Mashida et al. *Chem Pharm Bull Jpn* **26** 596 1978, Schuldiner et al. *Eur J Biochem* **25** 64 1972,]

**Acridone** [578-95-0] **M 195.2, m >300°, pK<sub>1</sub> -0.32 (basic), pK<sub>2</sub> 12.80 (14.0, acidic).** Dissolve ~1g in ca 1% NaOH (100ml), add 3M HCl to pH 4 when acridone separates as a pale yellow solid with **m** just above 350° (sharp). It can be recrystallised from large volumes of H<sub>2</sub>O to give a few mg. It is soluble in 160 parts of boiling EtOH (540 parts at 22°) [Albert & Phillips *J Chem Soc* 1294 1956]. A few decigrams are best crystallised as the *hydrochloride* from 400 parts of 10N HCl (90% recovery) from which the free base is obtained by washing the salt with H<sub>2</sub>O. A small quantity can be recrystallised (as the neutral species) from boiling AcOH. Larger quantities are best recrystallised from a mixture of 5 parts of freshly distilled aniline and 12.5 parts of glacial acetic acid. Acridone distils unchanged at atmospheric pressure, but the boiling point was not recorded, and some sublimation occurs below 350°. It has UV with λ<sub>max</sub> at 399nm. [see Albert, *The Acridines* Arnold Press pp 201, 372 1966; for pKa see Kalatzis *J Chem Soc (B)* 96 1969, *Beilstein* **23/9** V 7.]

**Acriflavine** [8048-52-0] **M 196.2, pK<sup>25</sup> >12.** Treat acriflavin twice with freshly precipitated AgOH to remove proflavine, then recrystallise it from absolute methanol [Wen & Hsu *J Phys Chem* **66** 1353 1962]. [*Beilstein* **22** III/IV 218.]

**Acriflavin Mixture (Euflavin, 3,6-diamino-10-methylacridinium chloride)** [8063-24-9] **M 259.7, m 179-181°.** Purify it by dissolving in 50 parts of H<sub>2</sub>O, shaking with a small excess of freshly precipitated and washed Ag<sub>2</sub>O. The mixture is set aside overnight at 0° and filtered. The cake is not washed. The pH of the filtrate is adjusted to 7.0 with HCl and evaporated to dryness. The residue is then crystallised twice from MeOH, twice from H<sub>2</sub>O and dried at 120°. Its UV spectrum has λ<sub>max</sub> at 452nm and a log<sub>e</sub> value of 4.67. It is a red powder which readily absorbs H<sub>2</sub>O. The solubility is increased in the presence of proflavin. The *dihydrochloride* is a deep red crystalline powder. It is available as a mixture of 3,6-diaminoacridinium chloride (35%) and its 10-metho-chloride (65%). [See Albert, *The Acridines* Arnold Press p 346 1966, Benda *Chem Ber* **45** 1787 1912]. [*Beilstein* **23** I 650.]

**Adenine (6-aminopurine, vitamin B<sub>4</sub>)** [73-24-5] **M 135.1, m 360-365°(dec rapid heating), pK<sub>1</sub><sup>25</sup> 4.12, pK<sub>2</sub><sup>25</sup> 9.83.** Crystallise adenine from distilled water. [*Beilstein* **26** III/IV 3561.]

**Adenosine** [58-61-7] **M 267.3, m 235-236°, [α]<sub>D</sub><sup>24</sup> -85° (c 2, 5% NaOH), [α]<sub>D</sub><sup>24</sup> -61.2° (c 1, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 3.48, pK<sub>2</sub><sup>25</sup> 12.5.** Crystallise adenosine from distilled water and dry it at 110°. It has been purified *via* the *picrate*, where ethanolic picric acid is added to adenosine and the *picrate* is filtered off and recrystallised from EtOH. It has **m** 180-185°(dec). Adenosine is recovered by dissolving 0.4g of the *picrate* in 80ml of hot H<sub>2</sub>O, treated with a small quantity of Dowex 1 anion exchange resin in the chloride form, and the resin is filtered off. The filtrate is treated with more resin and filtered again. One equivalent of aqueous NaOH is added to the colourless filtrate which is evaporated to 4ml and cooled to give 0.176g of adenosine **m** 236°. [Davoll et al. *J Chem Soc* 967 1948, Davoll & Lowy *J Am Chem Soc* **73** 1650 1951, *Beilstein* **26** III/IV 3598.]

**Adrenochrome (3-hydroxy-1-methyl-5,6-indoline-dione)** [54-06-8] **M 179.2, m 125-130° (dec).** It was

crystallised from MeOH/formic acid, as red crystals of the *hemihydrate*, and stored in a vacuum desiccator. The *mono-semicarbazone* (*Carbazochrome*) [69-81-8] M 236.2, crystallises as orange-red crystals from dilute EtOH with *m* ~203°(dec) and is haemostatic. [Heacock *Chem Rev* **59** 181 1959, *Beilstein* **21** III/IV 6434.]

**Aetioporphyrin I (2,4,6,8-tetraethyl-1,3,5,7-tetramethylporphyrin)** [448-71-5] M 478.7, *m* 360-363°, *pK*<sup>25</sup> 18. Purify it by chromatography on an Al<sub>2</sub>O<sub>3</sub> column (300g/300mg of porphyrin) and elute with CH<sub>2</sub>Cl<sub>2</sub>, evaporate the eluate and crystallise the residue from CH<sub>2</sub>Cl<sub>2</sub>/MeOH, pyridine or CHCl<sub>3</sub>/petroleum ether (purple prisms, *m* > 300°) [Smith *J Chem Soc Perkin Trans 1* 1471 1972]. The copper salt crystallises as red needles from pyridine/AcOH. It complexes with metals. The *dihydrobromide* [69150-58-9] M 640.5 separates from aetioporphyrin I in Et<sub>2</sub>O on addition of 10% of aqueous HBr after 2 days [Fischer & Treibs *Justus Liebigs Ann Chem* **457** 241 1927, Treibs & Dieter *Justus Liebigs Ann Chem* **513** 88-91 1934]. [*Beilstein* **26** III/IV 1915.]

**Agroclavin** [548-42-5] M 238.3, *m* 198-203°(dec), 205-206°, 210-212°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -155° (c 1, CHCl<sub>3</sub>), *pK*<sub>Est</sub> ~8.0. It crystallises from diethyl ether or Me<sub>2</sub>CO. The *hydrochloride* crystallises from H<sub>2</sub>O and has *m* 265-266°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -110° (CHCl<sub>3</sub>). [Plieninger et al. *Justus Liebigs Ann Chem* **743** 95 1971, *Beilstein* **23** III/IV 1623.]

**Ajmalicine ( $\delta$ -yohimbine, Raubacine)** [483-04-5] M 352.4, *m* 250-252°(dec), 257°(dec), 261-263°(dec), [ $\alpha$ ]<sub>546</sub> -76° (c 0.5, CHCl<sub>3</sub>), [ $\alpha$ ]<sub>D</sub><sup>20</sup> -45° (c 0.5, pyridine), -39° (c 0.25, MeOH), *pK*<sup>25</sup> 5.7(90% aqueous 2-methoxyethanol), 6.46(66% aqueous Me<sub>2</sub>NCHO). It crystallises from MeOH, EtOH or EtOAc. [*Beilstein* **27** III/IV 7927.]

**Ajmalicine hydrochloride** [4373-34-6] M 388.9, *m* 290°(dec), 294-295°(dec), [ $\alpha$ ]<sub>D</sub> -18.5° (c 0.5, MeOH). This hydrochloride crystallises from MeOH or EtOH. Its solubility is 0.315g/100g of H<sub>2</sub>O at 100°. [*Beilstein* **27** III/IV 7928.]

**Ajmaline ( $\gamma$ -yohimbine)** [4360-12-7] M 326.4, *m* 160° (MeOH), 205-206° (anhyd), [ $\alpha$ ]<sub>D</sub><sup>20</sup> +144° (c 0.8, CHCl<sub>3</sub>), *pK*<sup>25</sup> 8.15(80% aqueous 2-methoxyethanol), 8.4(60%aqueous Me<sub>2</sub>NCHO). Ajmaline crystallises from MeOH or EtOAc containing a little H<sub>2</sub>O to give the *trihydrate* *m* 158-160°. This loses 1H<sub>2</sub>O at 110° and all H<sub>2</sub>O at 150°. [*Beilstein* **23** III/IV 3212.]

**Ajmaline hydrochloride** [4410-48-4] M 388.9, *m* 140°. It crystallises from water as the *dihydrate* *m* 140°, and the *anhydrous salt* has *m* 252-255°, and [ $\alpha$ ]<sub>D</sub><sup>40</sup> +84.6° (c 1, H<sub>2</sub>O). The *dihydrochloride* crystallises from EtOH/Et<sub>2</sub>O with *m* 305-306°. [*Beilstein* **23** III/IV 3212.]

**Alizarin (1,2-dihydroxyanthraquinone)** [72-48-0] M 240.2, *m* 290°, *d*<sub>4</sub><sup>20</sup> 0.884, CI 58000, *pK*<sub>1</sub><sup>25</sup> 7.45, *pK*<sub>2</sub><sup>25</sup> 11.80. Alizarin crystallises from glacial acetic acid or 95% EtOH. It can also be sublimed at 110°/2mm. It is an indicator with  $\lambda_{\text{max}}$  at 452nm (pH 5.8) and 520nm (pH 7.2). [*Beilstein* **8** IV 3256.]

**Alizarin-3-methyliminodiacetic acid (Alizarin Complexone) (2H<sub>2</sub>O)** [3952-78-1] M 421.4, *m* 189°(dec), *pK*<sub>Est(1)}</sub> ~4.9, *pK*<sub>Est(2)}</sub> ~7.5. It is purified by suspending it in 0.1M NaOH (1g in 50ml), filtering the solution and extracting alizarin with 5 successive portions of CH<sub>2</sub>Cl<sub>2</sub>. Then add HCl dropwise to precipitate the reagent, stirring the solution in an ice bath. Filter the precipitate onto a glass filter, wash it with cold water and dry it in a vacuum desiccator over KOH [Ingman *Talanta* **20** 135 1973, *Beilstein* **14** IV 931].

**Alizarin Yellow R [5-(4-nitrophenylazosalicylic acid), Mordant Orange I]** [2243-76-7] M 287.2, *m* 253-254°(dec), >300°, CI 14030, *pK*<sup>25</sup> 11.17. The free acid is precipitated by adding HCl to an aqueous solution of the Na salt. After 2 recrystallisations from aqueous AcOH, it has *m* 255°(dec); [*m* 253-254°(dec) was reported by Hewitt & Fox *J Chem Soc* **79** 49 1901]. The free acid recrystallises from dilute AcOH as orange brown needles. The Na salt changes colour from yellow to red when the pH is increased from 10.2 to 12.0. [Woodland et al. *J Am Chem Soc* **75** 5835 1953, *Beilstein* **16** IV 372.]

**RS-Allantoin** [97-59-6] M 158.1, *m* 238°(dec). It crystallises from water or EtOH [Hartman et al. *Org Synth Coll Vol II* 21 1943]. [*Beilstein* **25** III/IV 4071.]

**Alloxan [2,4,5,6(1*H*,3*H*)pyrimidine, tetrone]** [50-71-5] **M 142.0, m ~170°(dec), pK<sup>25</sup> 6.64.** Crystallisation from water gives the *tetrahydrate*. *Anhydrous* crystals are obtained by crystallisation from acetone, glacial acetic acid or by sublimation *in vacuo*. [See below and *Beilstein* **24** H 500, **24** I 428, **24** II 301, **24** III/IV 2137.]

**Alloxan monohydrate** [2244-11-3] **M 160.1, m 255°(dec), pK<sup>25</sup> 6.64.** Recrystallisation from H<sub>2</sub>O gave the *tetrahydrate* in large prisms or rhombs. On heating at 100°, or on exposure to air, this is converted to the *monohydrate*. Dissolve it in its own weight of boiling H<sub>2</sub>O and cool it for several days below 0°; the *tetrahydrate* crystallises from solution much more slowly when free from HNO<sub>3</sub>. It is less soluble in bicarbonate solutions than in H<sub>2</sub>O. Drying the solid over H<sub>2</sub>SO<sub>4</sub> yields the *monohydrate*. The *anhydrous* crystals can be obtained by recrystallisation from dry Me<sub>2</sub>CO or AcOH followed by washing with dry Et<sub>2</sub>O, or by sublimation in a vacuum. On heating it turns pink at 230° and decomposes at *ca* 256°. It is acidic to litmus. [Hartman & Sheppard *Org Synth Coll Vol III* 37 1955.] It forms a compound with urea which crystallises from H<sub>2</sub>O in yellow needles that become red at 170° and decompose at 185-186°. [*Beilstein* **24** H 500, **24** I 428, **24** II 301, **24** III/IV 2137.]

**Alloxantin** [76-24-4] **M 286.2, m 253-255°(dec) (yellow at 225°).** Alloxantin crystallises from water or EtOH and is kept under nitrogen. It turns red in air. [*Beilstein* **26** III/IV 2782.]

**1-Allyl-6-amino-3-ethyluracil (Aminometradine)** [642-44-4] **M 195.2, m 143-144° (anhydrous).** It crystallises from water (as monohydrate). It is a diuretic. [*Beilstein* **24** III/IV 4133.]

**1-*N*-Allyl-3-hydroxymorphinan (Levallorphan)** [152-02-3] **M 283.4, m 180-182°, [α]<sub>D</sub><sup>20</sup> -89° (c 3, MeOH).** It crystallises from aqueous EtOH. It is a narcotic antagonist. [Schneider & Grüssner *Helv Chim Acta* **34** 2211 1951, Hellerbach *Helv Chim Acta* **39** 429 1956.]

**5-Allyl-5-isobutylbarbituric acid** [77-26-9] **M 224.3, m 139°, 139-140°, 140-142°, pK<sup>18</sup> 12.36.** It can be recrystallised from H<sub>2</sub>O or dilute EtOH, and sublimes at 100-120°/8-12mm. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, cyclohexane, tetralin and petroleum ether at 20°. [Butler et al. *J Am Chem Soc* **77** 1486 1955, *Beilstein* **24** III/IV 2006.]

**9-Aminoacridine [9-acridineamine]** [90-45-9] **M 194.2, m 241°, pK<sup>20</sup> 9.95.** It crystallises from EtOH or acetone and sublimes at 170-180°/0.04mm [Albert & Ritchie *Org Synth Coll Vol III* 53 1955, for hydrochloride, see below]. [*Beilstein* **22** H 463, **21** II 280, **21** III/IV 4174.]

**9-Aminoacridine hydrochloride monohydrate (Acramine yellow, Monacrin)** [52417-22-8] **M 248.7, m >355°, pK<sub>1</sub><sup>20</sup> 4.7, pK<sub>2</sub><sup>20</sup> 9.99.** Recrystallise it from boiling H<sub>2</sub>O (charcoal; 1g in 300 ml) to give pale yellow crystals with a neutral reaction. It is one of the most fluorescent substances known. At 1:1000 dilution in H<sub>2</sub>O it is pale yellow with only a faint fluorescence, but at 1:100,000 dilution it is colourless with an intense blue fluorescence. [Albert & Ritchie *Org Synth Coll Vol III* 53 1955; Falk & Thomas *Pharm J* **153** 158 1944, *Beilstein* **22** H 463, **21** II 280, **21** III/IV 4174.] See previous entry for the free base.

**2-Amino-4-anilino-*s*-triazine (Amanozine)** [537-17-7] **M 168.2, m 235-236°, pK<sub>Est</sub> ~5.5.** It crystallises from dioxane or 50% aqueous EtOH. [*Beilstein* **26** III/IV 1195.]

**4-Aminoantipyrene** [83-07-8] **M 203.3, m 109°.** It crystallises from EtOH or EtOH/ether. [*Beilstein* **25** III/IV 3554.]

**2-Aminobenzothiazole** [136-95-8] **M 150.2, m 132°, pK<sup>20</sup> 4.48.** The thiazole crystallises from H<sub>2</sub>O, aqueous EtOH, \*C<sub>6</sub>H<sub>6</sub> or petroleum ether. The *hydrochloride* crystallises from dilute HCl and has **m 240.5°**. [*Beilstein* **27** H 182, **27** III/IV 4824.]

**6-Aminobenzothiazole** [533-30-2] **M 150.2, m 87°, pK<sub>Est</sub> ~3.** It crystallises from aqueous EtOH, petroleum ether or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. The *hydrochloride* has **m 305°(dec)** from dilute HCl, and the *picrate* has **m 185°(dec)** from Me<sub>2</sub>CO. [Boggust & Cocker *J Chem Soc* 360 1949, *Beilstein* **27** III/IV 4884.]

**3-*o*-Aminobenzyl-4-methylthiazolium chloride hydrochloride** [534-94-1] **M 277.4, m 213°(dec)**. The hydrochloride crystallises from aqueous EtOH, and the iodide hydroiodide has **m 273°** (from aqueous HI). [Beilstein 27 III/IV 973.]

**4-Amino-1-benzylpiperidine** [50541-93-0] **M 190.3, b ~180°/20mm,  $d_4^{20}$  0.933,  $n_D^{20}$  1.543,  $pK_{Est(1)}$  ~8.3,  $pK_{Est(2)}$  ~10.4**. Purify it by distillation *in vacuo* and store it under N<sub>2</sub> because it absorbs CO<sub>2</sub>. The dihydrochloride salt [1205-72-7] has **m 270-273° (255°)** after recrystallisation from MeOH/EtOAc or EtOH. [Brookes *J Chem Soc* 3165, 3172 1957.] The 4-methylamino-1-benzylpiperidine derivative has **b 168-172°/17mm,  $n_D^{20}$  1.5367** [Reitsema & Hunter *J Am Chem Soc* 70 4009 1948]. The 1-(1-benzyl-4-piperidinyl)-3-cyano-2-methylisothiourea derivative has **m 160°** from CHCl<sub>3</sub>/Et<sub>2</sub>O [Preparation, IR, NMR: Ried et al. *Chem Ber* 116 1547 1983, Beilstein 22 III/IV 3752].

**2-Amino-4-chloro-6-methylpyrimidine** [5600-21-5] **M 143.6, m 184-186°,  $pK_{Est}$  ~1.0**. Recrystallise it from EtOH. [Beilstein 24 H 84.]

**2-Amino-5-chloropyridine** [1072-98-6] **M 128.6, m 135-136°,  $pK^{25}$  4.38**. Recrystallise this base from petroleum ether. It sublimes at 50°/0.5mm. [Beilstein 22 II 332, 22/8 V 541.]

**2-Amino-4-chloropyrimidine** [3993-78-0] **M 129.55, m 168-169°, m 170°,  $pK_{Est}$  ~1.2**. The pyrimidine crystallises in glistening plates from EtOH (**m 170°**, sintering at 167°). It has also been purified by sublimation in a vacuum and recrystallisation from H<sub>2</sub>O. [Hilbert & Johnson *J Am Chem Soc* 52 1155 1930, Beilstein 24 H 80, 25 III/IV 2117.]

**2-Amino-3,5-dibromopyridine** [35486-42-1] **M 251.9, m 103-104°,  $pK_{Est}$  ~2.4**. Steam distil it and recrystallise it from aqueous EtOH or petroleum ether. [Beilstein 22 H 431, 22 II 333, 22 III/IV 4041.]

**3-Amino-2,6-dichloropyridine** [62476-56-6] **M 164.0, m 119°, b 110°/0.3mm,  $pK_{Est}$  ~2.0**. Recrystallise it from water. [Beilstein 22 III/IV 4093.]

**2-Amino-4,6-dimethylpyridine** [5407-87-4] **M 122.2, m 69-70.5°,  $pK^{25}$  7.84**. Recrystallise this base from hexane, ether/petroleum ether or \*benzene. Residual \*benzene is removed over paraffin-wax chips in an evacuated desiccator. The dipicrate crystallises from EtOH and has **m 205-207°(dec)**. [Beilstein 22 III/IV 4210.]

**2-Amino-4,6-dimethylpyrimidine** [767-15-7] **M 123.2, m 152-153°,  $pK^{25}$  4.95**. Recrystallisation from water gives the pyrimidine with **m 197°**, but recrystallisation from acetone gives **m 153°**. [Beilstein 25 III/IV 2205.]

**2-(Aminomethyl)piperidine** [22990-77-8] **M 114.2, b 66-67°/12mm, 80-81°/18mm,  $d_4^{20}$  0.9406,  $n_D^{20}$  1.4854,  $pK_1^{20}$  6.33,  $pK_2^{20}$  9.70**. Dry (over Na<sub>2</sub>SO<sub>4</sub>) and distil the piperidine under vacuum from KOH. It has been purified *via* the Reinecate salt (**m 173-174°**) and its dipicrate salt (**m 201°**, from H<sub>2</sub>O). [Norton et al. *J Am Chem Soc* 68 1330 1946, Mortimer *Aust J Chem* 11 84 1958, Augustine *J Am Chem Soc* 81 4667 1959, Beilstein 22 III/IV 3765.]

**4-Amino-3-hydrazino-5-mercapto-1,2,4-triazole (Purpald)** [1750-12-5] **M 146.2, m 228-230°(dec), 234-235°(dec),  $pK_{Est(1)}$  ~2,  $pK_{Est(2)}$  ~3 (NH<sub>2</sub>),  $pK_{Est(3)}$  ~8 (SH)**. Recrystallise Purpald from H<sub>2</sub>O (0.6g in 300-400ml). The benzylidene derivative has **m 245-246°(dec)** from *i*-PrOH [Hoggarth *J Chem Soc* 4817 1952, Dickinson & Jacobson *J Chem Soc, Perkin Trans I* 975 1975. [Beilstein 26 III/IV 547.]

**5-Amino-8-hydroxyquinoline hydrochloride** [3881-33-2] **M 196.7,  $pK_1^{20}$  5.67,  $pK_2^{20}$  11.24**. Dissolve the hydrochloride in the minimum volume of MeOH, then add Et<sub>2</sub>O to initiate crystallisation. The crystals are filtered off and dried [Lovell et al. *J Phys Chem* 88 1885 1984]. The dihydrochloride [21302-43-2] has **M 233.1, m 279°(dec)**. [Beilstein 22 III/IV 5866.]

**4-Aminoimidazole-5-carboxamide hydrochloride (AICAR HCl)** [72-40-2] **M 162.6, m 255-256°(dec), pK<sub>Est(1)</sub> ~3.5, pK<sub>Est(2)</sub> ~9.4.** Recrystallise the hydrochloride from EtOH. [Kuroda & Hakko *J Heterocycl Chem* **30** 593 1993, Alhede et al. *J Org Chem* **56** 2139 1991, Cheru et al. *Heterocycles* **24** 1133 1992, *Beilstein* **25** II 221, **25** III/IV 4329.]

**6-Aminoindazole** [6967-12-0] **M 133.2, m 210°, pK<sup>25</sup> 3.99.** It is recrystallised from H<sub>2</sub>O or EtOH and sublimes in a vacuum. [*Beilstein* **25** H 317.]

**2-Amino-3-iodopyridine** [104830-06-0] **M 220.1, m 78-91°, 90-91.5°, pK<sub>Est</sub> ~4.9.** Purify this pyridine by recrystallisation from hexane. The *N-Me* derivative [113975-23-8] has **m 50°**, and distils at **b 129°/14mm**. [Sakamoto et al. *Chem Pharm Bull, Japan* **33** 4764 1985, Estel et al. *J Org Chem* **53** 2740 1988.]

**2-Amino-4-iodopyridine** [552331-00-7] **M 220.1, m 163-164°, pK<sub>Est</sub> ~5.1.** Purify this pyridine by recrystallisation from H<sub>2</sub>O. The *picrate* has **m 253-254°** (from H<sub>2</sub>O), the *N-acetyl* derivative has **m 150°** (from H<sub>2</sub>O), and the *N-benzoyl* derivative has **m 167-168°** (from aqueous EtOH). [Graf *Chem Ber* **64** 21 (25) 1931.]

**2-Amino-5-iodopyridine** [20511-12-0] **M 220.1, m 128-131°, 129-130°, 130°, pK<sub>Est</sub> ~4.5.** The pyridine can be purified by steam distillation. Separate the solid from the cooled distillate by filtration, acidify the filtrate, decolorise it with charcoal, concentrate it to ~200ml, make alkaline with KOH and cool. Filter the solid, add it to the original solid that was collected, and dry it *in vacuo*. It crystallises from \*C<sub>6</sub>H<sub>6</sub> in white needles. The *picrate* separates as yellow needles from hot EtOH or Me<sub>2</sub>CO, **m 240°**. [Caldwell et al. *J Am Chem Soc* **66** 1479 1944.]

**5-Amino-2-iodopyridine** [29958-06-0] **M 220.1, m 132°, pK<sub>Est</sub> ~2.6.** Purify it by recrystallisation from EtOH (white needles). [Caldwell et al. *J Am Chem Soc* **66** 1479 1944.]

**3-Amino-5-mercapto-1,2,4-triazole** [16691-43-3] **M 116.1, m 298°, pK<sub>Est(1)</sub> ~ 3.0, pK<sub>Est(2)</sub> ~ 9.** Recrystallise the triazole from H<sub>2</sub>O and dry it *in vacuo*. The *acetyl* derivative has **m 325°** (dec) after recrystallisation from H<sub>2</sub>O. [*Beilstein* **26** III/IV 1351.] It has also been recrystallised from EtOH/ H<sub>2</sub>O (3:1, 1g in 50 ml, 50% recovery), **m 300-302°** (dec subject to heating rate), ( $\lambda_{\max}$  263nm, log  $\epsilon$  4.12). The *S-benzyl* derivative, when crystallised from \*C<sub>6</sub>H<sub>6</sub>/EtOH (20:1), or CHCl<sub>3</sub>/Et<sub>2</sub>O has **m 109-111°** [Godfrey & Kruzer *J Chem Soc* 3437 1960, *Beilstein* **26** III/IV 1351.]

**2-Amino-4-methoxy-6-methylpyrimidine** [7749-47-5] **M 139.2, m 157-159°, 158-158.5°, 158-160°, pK<sub>Est</sub> ~6.0.** Recrystallise it from H<sub>2</sub>O. The *picrate* has **m 220-221°(dec)**. [Baker et al. *J Am Chem Soc* **69** 3072, 3075 1947, Sirakawa et al. *Yakugaku Zasshi* **73** 598 1953, Backer & Grevenstuk *Rec Trav Chim Pays Bas* **61** 291 1942, *Beilstein* **25** III/IV 3385.]

**8-Amino-6-methoxyquinoline** [90-52-8] **M 174.2, m 41-42°, 51°, b 137-138°/1mm, pK<sup>70.1</sup> 3.38.** Distil it under N<sub>2</sub> and at high vacuum, then recrystallise it several times from MeOH (0.4ml/g). It remains colourless for several months when purified in this way [Elderfield & Rubin *J Am Chem Soc* **75** 2963 1953]. The *hydrobromide* [312693-53-1] **M 255.1** has **m 238°(dec)**. [*Beilstein* **22** III/IV 5934.]

**7-Amino-4-methylcoumarin** [26093-31-2] **M 175.2, m 224-229°(dec), pK<sub>Est</sub> ~3.2.** Dissolve it in 5% HCl, filter and basify with 2M ammonia. The precipitate is dried in a vacuum and recrystallised from dilute EtOH. It yields a blue solution and is light sensitive. [Sigler *Synthesis* 614 1980, Kanaoka et al. *Chem Pharm Bull Jpn* **30** 1485 1982, *Beilstein* **18/11** V 445.]

**2-Amino-3-methylpyridine (2-amino-3-picoline)** [1603-40-3] **M 108.1, m 33.2°, b 221-222°, pK<sup>25</sup> 7.24.** Recrystallise the picoline three times from \*benzene, most of the residual \*benzene being removed from the crystals by standing over paraffin wax chips in an evacuated desiccator. The amine is also transferred to a separating funnel under N<sub>2</sub>, and left in contact with NaOH pellets for 3 hours with occasional shaking. It is then placed in a vacuum distilling flask where it is refluxed gently in a stream of dry N<sub>2</sub> before fractionally distilling it. [Mod et al. *J Phys Chem* **60** 1651 1956, *Beilstein* **22/9** V 212.]

**2-Amino-4-methylpyridine (2-amino-4-picoline)** [695-34-1] M 108.1, m 99.2°, b 230°, pK<sup>25</sup> 7.48. Crystallise it from EtOH or a 2:1 \*benzene/acetone mixture, and dry it under vacuum as in the previous entry. [Beilstein 22/9 V 325.]

**2-Amino-5-methylpyridine (6-amino-3-picoline)** [1603-41-4] M 108.1, m 76.5°, b 227°, pK<sup>25</sup> 7.22. Crystallise it from acetone. [Beilstein 22/9 V 289.]

**2-Amino-6-methylpyridine (6-amino-2-picoline)** [1824-81-3] M 108.1, m 44.2°, b 208-209°, pK<sup>25</sup> 7.41. Crystallise it three times from acetone and dry it under vacuum at ca 45°. Alternatively, keep it in contact with NaOH pellets for 3 hours, with occasional shaking, decant and fractionally distil it [Mod et al. *J Phys Chem* 60 1651 1956]. It also recrystallises from CH<sub>2</sub>Cl<sub>2</sub> on addition of petroleum ether. [Marzilli et al. *J Am Chem Soc* 108 4830 1986, Beilstein 22/9 V 210.]

**2-Amino-4-methylpyrimidine** [108-52-1] M 109.1, m 159-160°, 161°, pK<sup>20</sup> 4.15. Crystallise the pyrimidine from H<sub>2</sub>O or EtOH and sublime it in a vacuum. The *picrate* crystallises from EtOH and has m 235-236°(dec). [Beilstein 24 H 84, 25 III/IV 2152.]

**2-Amino-5-methylpyrimidine** [50840-23-8] M 109.1, m 193.5°, pK<sub>Est</sub> ~4.0. Crystallise it from water or \*benzene/petroleum ether and sublime it at 50°/0.5mm. [Beilstein 24 H 87.]

**4-Amino-2-methylquinoline (4-aminoquinoline)** [6628-04-2] M 158.2, m 168°, b 333°/760mm, pK<sup>20</sup> 9.42. Recrystallise it from \*benzene/petroleum ether. [Beilstein 22/10 V 347.]

**6-Aminonicotinic acid** [3167-49-5] M 138.1, m 312°(dec), pK<sub>Est(1)</sub> ~2.2 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~6.5. Crystallise the acid from aqueous acetic acid. Dry it *in vacuo* at 70°. [Beilstein 22 III/IV 6726.]

**6-Aminopenicillanic acid** [551-16-6] M 216.2, m 208-209°, [α]<sub>546</sub> +327° (in 0.1M HCl), pK<sub>1</sub><sup>25</sup> 2.30, pK<sub>2</sub><sup>25</sup> 4.90. This acid crystallises from water. [Kleppe & Stroninger *J Biol Chem* 254 4856 1979, Beilstein 27 III/IV 2858.]

**2-Aminoperimidine** [28832-64-6] M 183.1, m 239°, b 170-175°/1.5mm, pK<sub>Est</sub> ~7.9 (free base). It crystallises from EtOH/H<sub>2</sub>O (1:1). It precipitates as the *hydrochloride* with dilute HCl which has m 282°. [Dasgupta et al. *Anal Chim Acta* 94 205 1977, Beilstein 24 H 193, 25 III/IV 2677.]

**2-Aminoperimidine hydrobromide** [40835-96-9] M 264.1, m 299°, pK<sub>Est</sub> ~7.9 (free base). Purify the hydrobromide by boiling a saturated aqueous solution with charcoal, filtering and leaving the salt to crystallise. Store this *dihydrate salt* in a cool, stoppered flask in the dark place. The *anhydrous salt* is obtained by heating at 80°/4 hours, and it is hygroscopic. The solubilities of the hydrobromide at 26° are 2.4% in EtOH, 0.6% in H<sub>2</sub>O, 0.3% in Et<sub>2</sub>O, 0.1% in Me<sub>2</sub>CO and 0.003% in \*C<sub>6</sub>H<sub>6</sub>. [Dasgupta et al. *Anal Chim Acta* 94 205 1977, Dasgupta & West *Microchim Acta* 2 505 1978, Dasgupta et al. *Anal Chem* 50 1793 1978, Beilstein 24 H 193, 25 III/IV 2677.]

**2-Aminopyridine** [504-29-0] M 94.1, m 58°, b 204-210°, pK<sub>1</sub><sup>25</sup> -7.6, pK<sub>2</sub><sup>25</sup> 6.71. It crystallises from \*benzene/petroleum ether (b 40-60°) or CHCl<sub>3</sub>/petroleum ether. [Beilstein 22/8 V 280.]

**3-Aminopyridine** [462-08-8] M 94.1, m 64°, b 248°, pK<sub>1</sub><sup>25</sup> -1.5, pK<sub>2</sub><sup>25</sup> 6.03. It crystallises from \*benzene, CHCl<sub>3</sub>/petroleum ether (b 60-70°), or \*benzene/petroleum ether (4:1). [Beilstein 22/9 V 3.]

**4-Aminopyridine** [504-24-5] M 94.1, m 160°, b 180°/12-13mm, pK<sub>1</sub><sup>25</sup> -6.55, pK<sub>2</sub><sup>25</sup> 9.11 (9.18). Crystallise the aminopyridine from \*benzene/EtOH, then recrystallise it twice from water, then crush and dry it for 4 hours at 105° [Bates & Hetzer *J Res Nat Bur Stand* 64A 427 1960]. It has also been crystallised from EtOH, \*benzene, \*benzene/petroleum ether, toluene and sublimes in a vacuum. [Beilstein 22/9 V 106.]

**2-Aminopyrimidine** [109-12-6] **M 95.1, m 126-127.5°, pK<sup>20</sup> 3.45.** Crystallise 2-aminopyrimidine from \*C<sub>6</sub>H<sub>6</sub>, EtOH or H<sub>2</sub>O. [Beilstein 25 III/IV 2071.]

**4-Aminopyrimidine** [591-54-8] **M 95.1, m 149-151°, 154-156°, pK<sup>25</sup> 5.69.** Recrystallise 10.6g of aminopyrimidine from hot EtOAc (200ml) to give 7.4g colourless needles as first crop; evaporation to 25ml gives a second crop of 1.7g. The *hydroiodide* has **m 180°**. The *picrate* has **m 225°**. [Brown *J Soc Chem Ind* (London) 69 353 1950, Beilstein 24 H 81, 24 III/IV 2130.]

**5-Aminopyrimidine** [591-55-9] **M 95.1, m 171-172° (with sublimation), pK<sup>25</sup> 2.52.** It is purified by conversion to the MgCl<sub>2</sub> complex in a small volume of H<sub>2</sub>O. The complex (~ 5g) is dissolved in the minimum volume of hot H<sub>2</sub>O, passed through a column of activated Al<sub>2</sub>O<sub>3</sub> (200g), and the column is washed with EtOH. Evaporation of the EtOH gives a colourless residue of the aminopyrimidine which is recrystallised from \*C<sub>6</sub>H<sub>6</sub> (toluene could also be used) which forms needles at first, then prisms. It melts with sublimation. Acetylation yields *5-acetamidopyrimidine* which crystallises from \*C<sub>6</sub>H<sub>6</sub>, **m 148-149°**. [Whittaker *J Chem Soc* 1565 1951.]

**Aminopyrine (4-dimethylaminoantipyrene)** [58-15-1] **M 231.3, m 107-109°, 108°, pK<sub>1</sub><sup>25</sup> -2.22, pK<sub>2</sub><sup>25</sup> 4.94.** It crystallises from petroleum ether, sublimes between 80° and 90°, and forms metal complexes. [Beilstein 25 H 452, 25 III/IV 3555.]

**3-Aminoquinoline** [580-17-6] **M 144.2, m 93.5°, pK<sub>1</sub><sup>20</sup> -0.58, pK<sub>2</sub><sup>20</sup> 4.94.** It crystallises from \*C<sub>6</sub>H<sub>6</sub>, toluene, hexane and aqueous EtOH. [Beilstein 22 III/IV 4605, 22/10 V 233.]

**4-Aminoquinoline** [578-68-7] **M 144.2, m 155-155.5°, 158°, pK<sub>1</sub><sup>25</sup> -7.11(5.99), pK<sub>2</sub><sup>20</sup> 9.13.** It has been purified by zone refining and recrystallisation from \*C<sub>6</sub>H<sub>6</sub>, EtOH or H<sub>2</sub>O. The *hydrochloride* has **m 308°** (from MeOH), and the *picrate* has **m 277°** (from EtOH). [Albert et al. *J Chem Soc* 2240 1948, Beilstein 22 III/IV 4611, 22/10 V 341.]

**5-Aminoquinoline** [611-34-7] **M 144.2, m 110°, b 184°/10mm, 310°/760mm, pK<sub>1</sub><sup>20</sup> 0.97(0.49), pK<sub>2</sub><sup>20</sup> 5.42.** It crystallises from pentane and from \*benzene or EtOH. The *picrate* has **m 209-210°(dec)** (202° dec) (from aqueous EtOH). [Beilstein 22 III/IV 4669, 22/10 V 297.]

**6-Aminoquinoline** [580-15-4] **M 144.2, m 117-119°, 120°, b 146°/0.3mm, 192-195°/14mm, pK<sub>1</sub><sup>20</sup> 1.63, pK<sub>2</sub><sup>20</sup> 5.59.** It is purified by column chromatography on a SiO<sub>2</sub> column using CHCl<sub>3</sub>/MeOH (4:1) as eluent. It crystallises from \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and is an **irritant**. The *stypnate* has **m 239-240°** (from EtOH) and **m 242-243°** (from aqueous Me<sub>2</sub>CO). [Barrett et al. *J Chem Soc* 50, 57 1953, Beilstein 22 III/IV 4681, 22/10 V 303.]

**8-Aminoquinoline** [578-66-5] **M 144.2, m 70°, b 140.5-141°/7mm, 123°/5mm, pK<sub>1</sub><sup>20</sup> -0.52, pK<sub>2</sub><sup>20</sup> 3.95.** 8-Aminoquinoline crystallises from EtOH, ligroin, octane or H<sub>2</sub>O, and complexes with metals. [Beilstein 22 III/IV 4708, 22/10 V 316.]

**2-Aminothiazole** [96-50-4] **M 108.1, m 93°, b 140°/11mm, pK<sup>20</sup> 5.36.** It crystallises from petroleum ether (b 100-120°), or EtOH. [Beilstein 27 III/IV 4574.]

**1-Amino-1,2,4-triazole** [24994-60-3] **M 84.1, m 91-93°, pK<sub>Est</sub> ~2.** The triazole crystallises from water. [Barszez et al. *J Chem Soc, Dalton Trans* 2025 1986, Temple & Montgomery *1,2,4-Triazoles — The Chemistry of Heterocyclic Compounds* Vol 37 (Weissberger & Taylor eds.). Wiley & Sons NY 1981, ISBN 0-471-0656-6.]

**3-Amino-1,2,4-triazole (Amitrol)** [61-82-5] **M 84.1, m 159°, pK<sub>1</sub><sup>20</sup> 4.04, pK<sub>2</sub><sup>20</sup> 11.08.** It crystallises from EtOH (charcoal), then three times from dioxane [Williams et al. *J Phys Chem* 61 261 1957]. [Beilstein 26 H 137.] **Possible carcinogen.** [Beilstein 26 H 137, Temple & Montgomery *1,2,4-Triazoles — The Chemistry of Heterocyclic Compounds* Vol 37 (Weissberger & Taylor eds.). Wiley & Sons NY 1981, ISBN 0-471-0656-6.]

**4-Amino-1,2,4(4H)-triazole** [584-13-4] **M 84.1, m 80-81°, pK<sup>25</sup> 3.23.** It crystallises from EtOH/ Et<sub>2</sub>O or H<sub>2</sub>O. The *hydrochloride* has **m 151-152°** (from EtOH, 1g/10ml). [Allen & Bell *Org Synth Coll Vol III* 96 1955, Barszez et al. *J Chem Soc, Dalton Trans* 2025 1986, *Beilstein* 26 H 16, 26 II 7, 26 III/IV 40, Temple & Montgomery *1,2,4-Triazoles —The Chemistry of Heterocyclic Compounds Vol 37* (Weissberger & Taylor eds.). Wiley & Sons NY 1981, ISBN 0-471-0656-6.]

**7-Amino-4-(trifluoromethyl)coumarin**, [53518-15-3] **M 229.1, m 222°, pK<sub>Est</sub> ~3.1.** Purify the coumarin by column chromatography on a C18 column, elute with acetonitrile/0.01M H<sub>2</sub>O/HCl (1:1), and recrystallise it from isopropanol. *Alternatively*, it is eluted from a silica gel column with CH<sub>2</sub>Cl<sub>2</sub>, or by extracting a CH<sub>2</sub>Cl<sub>2</sub> solution (4g/L) with 1M aqueous NaOH (3 x 0.1L), followed by drying (MgSO<sub>4</sub>), filtration and evaporation. [Bissell *J Org Chem* 45 2283 1980, Zimmermann et al. *Anal Biochem* 70 258 1976.]

**4(6)-Aminouracil (4-amino-2,6-dihydropyrimidine)** [873-83-6] **M 127.1, m >350°, pK<sub>1</sub><sup>20</sup> 0.00 (basic), pK<sub>2</sub><sup>20</sup> 8.69 (acidic), pK<sub>3</sub><sup>20</sup> 15.32 (acidic).** Purify the aminouracil by dissolving it in 3M aqueous NH<sub>3</sub>, filtering hot, and adding 3M formic acid until precipitation is complete. Cool, filter off (or centrifuge), wash well with cold H<sub>2</sub>O, then EtOH and dry it in air. Dry it further in a vacuum at ~80°. [Barlin & Pfeiderer *J Chem Soc (B)* 1424 1971, *Beilstein* 25 III/IV 4107.]

**Amodiaquin [4-(3-dimethylaminomethyl-4-hydroxyanilino)-7-chloroquinoline]** [86-42-0] **M 355.9, m 208°(dec).** Amodiaquin crystallises from 2-ethoxyethanol or EtOH. [Burckhalter et al. *J Am Chem Soc* 70 1363 1948, *Beilstein* 22 III/IV 4647.]

**2-n-Amylpyridine (2-n-pentylpyridine)** [2294-76-0] **M 149.2, b 63.0°/2mm, 206.5-207°/~760mm, n<sub>D</sub><sup>26</sup> 1.4861, pK<sup>25</sup> 6.00.** Dry it with NaOH for several days, then distil it from CaO under reduced pressure, taking the middle fraction and redistilling it. The *picrate* has **m 72-72.8°** (from EtOH). [*Beilstein* 20 III/IV 2835.] *3-n-Pentylpyridine*, purified as the 2-isomer, has **b 110-112°/20mm, 224-226°/748mm**, and the *picrate* has **m 79.5-80°(from EtOH)**. [*Beilstein* 20 III/IV 2835.]

**4-n-Amylpyridine** [2961-50-4] **M 149.2, b 78.0°/2.5mm, 229-230°/~760mm, n<sub>D</sub><sup>20</sup> 1.4908, pK<sub>Est</sub> ~6.1.** It is dried with NaOH for several days, then distilled from CaO under reduced pressure, taking the middle fraction and redistilling it. The *picrate* has **m 104°** (from EtOH). [*Beilstein* 20 III/IV 2836.]

**Antipyrine [2,3-dihydro-1,5-dimethyl-3-oxo-2-phenylpyrazole]** [60-80-0] **M 188.2, m 114°, b 319°, pK<sup>25</sup> 1.45.** Antipyrine crystallises from EtOH/water mixture, \*benzene, \*benzene/petroleum ether or hot water (charcoal), and the crystals are dried under a vacuum. [*Beilstein* 24 H 27, 24 III/IV 75.]

**Aspergillie acid (2-hydroxy-3-isobutyl-6-[1-methylpropyl]pyrazine 1-oxide)** [490-02-8] **M 224.3, m 97-99°, pK<sup>25</sup> 5.5, [α]<sub>D</sub><sup>20</sup> +13.3° (c 4, EtOH).** It is recrystallised from MeOH and is sublimed at 80°/10<sup>-3</sup>mm. [Dutcher *J Biol Chem* 171 321 1947, *Beilstein* 24 III/IV 235.]

**Atropine [tropine (±)-tropate]** [51-55-8] **M 289.4, m 114-116°, pK<sub>1</sub><sup>20</sup> 4.35, pK<sup>18</sup> 9.85.** It crystallises from acetone or hot water, and sublimes at ~100°/high vacuum. Its solubility is 0.22% (~20°) and 1.3%(100°) in H<sub>2</sub>O, and 83%(60°). The *sulfate monohydrate* has **m 190-194°**. [*Beilstein* 21 IV 183, 21/I V 235.] *Antispasmodic.*

**8-Azaadenine [1123-54-2] M 136.1, m 345°(dec), pK<sub>1</sub><sup>20</sup> 2.65, pK<sub>2</sub><sup>20</sup> 6.29.** 8-Azaadenine crystallises from H<sub>2</sub>O. [*Beilstein* 25 III/IV 4157.]

**2-Azacyclotridecanone (lauro lactam)** [947-04-6] **M 197.3, m 152°.** 2-Azacyclotridecanone crystallises from CHCl<sub>3</sub> and is stored over P<sub>2</sub>O<sub>5</sub> in a vacuum desiccator. [*Beilstein* 26/I V 566.]

**8-Azaguanine [134-58-7] M 152.1, m >300°, pK<sub>1</sub><sup>20</sup> 1.04, pK<sub>2</sub><sup>20</sup> 6.29.** Dissolve it in hot M NH<sub>4</sub>OH, filter, and cool. Recrystallise it, and wash it with water, then dry it in a vacuum. [*Beilstein* 26 III/IV 4171.]

**7-Azaindole (1H-pyrrolo[2,3b]pyridine)** [271-63-6] **M 118.1, m 105-106°, pK<sup>20</sup> 4.57.** Recrystallise it

repeatedly from EtOH, then sublime it in a vacuum [Tokumura et al. *J Am Chem Soc* **109** 1346 1987]. The *N*-acetate has *m* 65-66° (from \*C<sub>6</sub>H<sub>6</sub>), and the *picrate* has *m* 232-233° (from Me<sub>2</sub>CO) [Clemo & Swan *J Chem Soc* 603 1945, *Beilstein* **23** III/IV 1105.]

**1-Azaindolizine (1,7a-diazaindene, imidazo[1,2-a]pyridine)** [274-76-0] **M 118.1, b 72-73°/1mm, pK<sup>20</sup> 1.43 (in aqueous HCl).** 1-Azaindolizine is purified by distillation or gas chromatography. [Bower & Ramage *J Chem Soc* 4506 1957, Armarego *J Chem Soc* 4226 1964, *Beilstein* **23** II 1554, **23** III/IV 1104.]

**8-Azapurine (1*H*-1,2,3-triazolo[4,5-*d*]pyrimidine, 1,2,3,4,6[3*H*]penta-azaindene)** [273-40-5] **M 121.1, m 174-175° (effervescence, m depends on heating rate), pK<sub>1</sub><sup>20</sup> 2.05 (equilib with covalent hydrate), pK<sub>2</sub><sup>20</sup> 4.84.** Sublime 8-azapurine at 120-130°/0.01mm and recrystallise it from 3 parts of EtOH. [Albert *J Chem Soc(B)* 427 1966, *Beilstein* **26** III/IV 4108.]

**Azetidine (trimethyleneimine)** [503-29-7] **M 57.1, b 19°/132.5mm, 61.3-61.5°/760mm, d<sub>4</sub><sup>20</sup> 0.846, n 1.432, pK<sup>25</sup> 11.29.** Azetidine is a flammable, hygroscopic liquid smelling of ammonia, which absorbs CO<sub>2</sub> from air and should be kept under Argon. Purify it by drying it over solid KOH and distilling it through a short Vigreux column (p 11) at atmospheric pressure (under Argon) and keeping the pot temperature below 210°. It is moisture sensitive. The *hydrochloride* [36520-39-5] **M 93.6** has *m* > 300° and the *hydroiodide* has *m* 146.5°(from EtOH). The *N*-Me derivative has *m* 112°(from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether), and the *N*-phenylcarbamoyl derivative has *m* 189-190°(from EtOH). [Searles et al. *J Am Chem Soc* **78** 4917 1956, *Beilstein* **20** H 2, **20** I 3, **20** II 3, **20** III/IV 53, **20** V 136.]

**Aziridine (ethyleneimine)** [151-56-4] **M 43.1, b 55-56°/756mm, 56°/760mm, d<sub>4</sub><sup>24</sup> 0.8321, pK<sup>25</sup> 8.00.** Redistil it in an Ar or N<sub>2</sub> atmosphere in a fume hood, and store it over KOH in sealed bottles in a refrigerator. Commercial aziridine has been dried over sodium and distilled from the metal through an efficient column before use [Jackson & Edwards *J Am Chem Soc* **83** 355 1961, Wenker *J Am Chem Soc* **57** 2328 1935]. It is a weaker base than Me<sub>2</sub>NH (pK<sup>25</sup> 10.87) but is caustic to the skin. It should not be inhaled, causes inflammation of the eyes, nose and throat, and one may become sensitised to it. It is soluble in H<sub>2</sub>O, has an ammoniacal smell and reacts with CO<sub>2</sub>. Pure aziridine is comparatively stable but polymerises in the presence of traces of H<sub>2</sub>O and is occasionally explosive in the presence of acids. CO<sub>2</sub> is sufficiently acidic to cause polymerisation (forms linear polymers) which is not free radical promoted. It is stable in the presence of bases. The violet 2:1 *Cu* complex crystallises from EtOH containing a few drops of aziridine and adding Et<sub>2</sub>O, and has *m* 142°(dec). The *picrate* has *m* 142°. [O'Rourke et al. *J Am Chem Soc* **78** 2159 1956.] It has also been dried over BaO and has been distilled from sodium under nitrogen. [Allen et al. *Org Synth Coll Vol IV* 433 1963, *Beilstein* **20** III/IV 1.] **TOXIC.**

**Azulen(1,2-*b*)thiophene** [25043-00-9] **M 184.2.** It is crystallised from cyclohexane, then sublimed *in vacuo*.

**Azulen(2,1-*b*)thiophene** [248-13-5] **M 184.2.** It is crystallised from cyclohexane, then sublimed *in vacuo*.

**Azure A (3-amino-7-dimethylaminophenazin-5-ium chloride)** [531-53-3] **M 291.8, CI 52005, m > 290°(dec), λ<sub>max</sub> 633nm, pK<sup>25</sup> 7.2.** Azure A has been twice recrystallised from H<sub>2</sub>O and dried at 100°/1 hour in an oven. [*Beilstein* **27** III/IV 5151.]

**Azure B (3-dimethylamino-7-methylaminophenazin-5-ium chloride)** [531-55-5] **M 305.8, CI 52010, m > 201°(dec), λ<sub>max</sub> 648nm, pK<sup>25</sup> 7.4.** Azure B has been twice recrystallised from H<sub>2</sub>O and dried at 100°/1 hour in an oven. [*Beilstein* **27** III/IV 5151.]

**Azure C (3-amino-7-methylaminophenazin-5-ium chloride)** [531-57-7] **M 277.8, λ<sub>max</sub> 616nm, pK<sup>25</sup> 7.0.** Azure C has been twice recrystallised from H<sub>2</sub>O, and dried at 100°/1 hour in an oven.

**Barbituric acid [6-hydroxypyrimidin-2,4-dione]** [67-52-7] **M 128.1, m 250°(dec), pK<sub>1</sub><sup>25</sup> 3.99, pK<sub>2</sub><sup>25</sup> 12.5.** Recrystallise it twice from H<sub>2</sub>O, then dry it for 2 days at 100°. [*Beilstein* **24** III/IV 1873.]

**Benzimidazole** [51-17-2] M 118.1, m 172-173°, pK<sub>1</sub><sup>25</sup> 5.53, pK<sub>2</sub><sup>25</sup> 11.70. It crystallises from water or aqueous EtOH (charcoal) and is dried at 100° for 12 hours. [Beilstein 23 H 131, 23/6 V 196.]

**2-Benzimidazolylacetonitrile** [4414-88-4] M 157.2, m 200-205°(dec), 209.7-210.7°(corrected), 210°. It is recrystallised from aqueous EtOH. It has also been recrystallised from hot H<sub>2</sub>O using charcoal, and finally from aqueous EtOH. [Copeland & Day *J Am Chem Soc* 65 1072 1943, Beilstein 25 III/IV 820.]

**Benzo-15-crown-5** [14098-44-3] M 268.3, m 78-80°. It is recrystallised from *n*-heptane. **IRRITANT.** [Vögtle ed, Host Guest Complex Chemistry in *Topics in Current Chemistry* p98 1981, Beilstein 19/10 V 618.]

**Benzo-18-crown-6** [14098-24-9] M 312.2, m 42-45°, 43-43.5°. Purify it by passage through a DEAE cellulose column in cyclohexane. It recrystallises from *n*-hexane. Its *thiourea complex* has m 127° [5-6 mol of urea to ether, Pedersen *J Org Chem* 36 1690 1971]. The stability constants of the Na<sup>+</sup>, K<sup>+</sup>, Rb<sup>+</sup>, Cs<sup>+</sup>, Tl<sup>+</sup> and Ba<sup>2+</sup> complexes are described in Hofmanova et al. *Inorg Chim Acta* 28 73 1978. [NMR: Live & Chan *J Am Chem Soc* 98 3769 1976]. [Beilstein 19/12 V 618.] **IRRITANT.**

**Benzo[3,4]cyclobuta[1,2-*b*]quinoxaline** [259-57-4] M 204.2, m dec >250°. It is purified by sublimation under reduced pressure.

**Benzofuran (coumarone)** [271-89-6] M 118.1, b 62-63°/15mm, 97.5-99.0°/80mm, 170-173°/atm, 173-175°(169)/760mm, d<sub>4</sub><sup>20</sup> 1.0945, n<sub>D</sub><sup>20</sup> 1.565. Benzofuran is steam distilled, dissolved in Et<sub>2</sub>O, washed with 5% aqueous NaOH, saturated NaCl, dried (Na<sub>2</sub>SO<sub>4</sub>), evaporated and redistilled. The UV has λ<sub>max</sub> at 245, 275, 282nm (log ε 4.08, 3.45, 3.48). The *picrate* has m 102-103°. [Burgstahler & Worden *Org Synth Coll Vol V* 251 1973, NMR: Black & Heffernan *Aust J Chem* 18 353 1965, Beilstein 17/2 V 3.]

**2-Benzofurancarboxylic acid** [496-41-3] M 162.1, m 192-193°, pK<sub>Est</sub> ~3.2. The acid crystallises from water. [Beilstein 18/6 V 419.]

**Benzofurazan** [273-09-6] M 120.1, m 55°. Purify benzofurazan by steam distillation from a dilute alkaline solution, crystallisation from EtOH (long white needles) and by sublimation. [Green & Rowe *J Chem Soc* 101 2456 1912, Green & Rowe *J Chem Soc* 111 618 1917, Ghosh et al. *J Med Chem* 15 255 1972, Beilstein 27 H 568, 27 I 573, 27 III/IV 7115.]

**Benzofuroxan (benzofurazan-1-oxide, benz[1,2,5]oxadiazol-1-oxide)** [480-96-6] M 136.1, m 70-71°, 72-73°. Purify the 1-oxide by dissolving 3.6g in 45ml of 95% EtOH + 15ml H<sub>2</sub>O, boil, filter hot and cool to 25°. The yellow crystalline solid is steam volatile (but less so than benzofurazan, above) and has a peculiar pungent odour. Its UV has λ<sub>max</sub> at 355nm (EtOH). It was used as a dehydrogenation oxidant (Pätzold et al. *Synth Commun* 22 281 1992). [Mallory *Org Synth Coll Vol IV* 74 1963, Boulton & Ghosh *Adv Heterocyclic Chem* 10 1 1969, Bo et al. *J Am Chem Soc* 79 1750 1977, Beilstein 27 I 622, 27 II 629, 27 III/IV 7115.]

**5,6-Benzoquinoline (benzo[*f*]quinoline)** [85-02-9] M 179.0, m 93°, 94°, b 350°, pK<sup>20</sup> 5.11. Purify as 3,4-benzoquinoline above. The *picrate* has m 258.1-259° (from EtOH or H<sub>2</sub>O). [Albert et al. *J Chem Soc* 2240 1948, Beilstein 20 III/IV 4009, 20/8 V 220.]

**7,8-Benzoquinoline (benzo[*h*]quinoline)** [230-27-3] M 179.0, m 52.0-52.5°, pK<sup>20</sup> 4.21. Purify it as for 3,4-benzoquinoline above. The *picrate* has m 196°(from Me<sub>2</sub>CO). [Beilstein 20 H 463, 20 III/IV 4003, 20/8 V 215.]

**1,2,3-Benzothiadiazole** [273-77-8] M 136.2, m 35°, 36-37°, b 63°/0.5mm, pK<sub>Est</sub> ~<0. 1,2,3-Benzothiadiazole crystallises from petroleum ether, and has λ<sub>max</sub> at 264 and 306nm in hexane. [Overberger et al. *J Org Soc* 24 1407 1959, Beilstein 27 III/IV 7113.]

**2,1,3(1,2,5)-Benzothiadiazole** [272-13-2] M 136.2, m 44°, b 206°/760mm, pK<sub>Est</sub> <0. 2,1,3-Benzothiadiazole crystallises from petroleum ether and has UV with λ<sub>max</sub> at 221-222, 304 and 330nm in EtOH. [Sanvicki & Carr

*J Org Soc* **22** 503 1959, *Beilstein* **27** III/IV 7118.]

**1-Benzothiophene** (benzo[b]thiophene, thianaphthene) [95-15-8] **M** 134.2, **m** 29-32°, 30°, 31-32°, 32°, **b** 100°/16mm, 103-105°/20mm, 221-222°/760mm,  $d_4^{32.2}$  1.1484,  $n_D^{39}$  1.6306. 1-Benzothiophene has the odour of naphthalene. If the IR spectrum is not very good, then suspend it in a faintly alkaline aqueous solution and steam distil it. Extract the distillate with Et<sub>2</sub>O, dry the extract (CaCl<sub>2</sub>), filter, evaporate the solvent and fractionate the residue. The distillate sets solid. The *sulfoxide* has **m** 142°, the *picrate* has **m** 148-149° (yellow crystals from EtOH) and the *stypnate* has **m** 136-137°. [Hansch & Lindwall *J Org Chem* **10**, 381 1945, Meyer & Meyer *Chem Ber* **52B** 1249 1919, Weisgerber & Kruber **53** 1551 1920, Iddon & Scrowston *Adv Heterocycl Chem* **11** 177 1970, *Beilstein* **17/2** V 6.]

**1,2,3-Benzotriazole** [95-14-7] **M** 119.1, **m** 96-97°, 98.5°, 100°, **b** 159°/0.2mm, 204°/15mm, **pK<sub>1</sub><sup>20</sup>** 1.6, **pK<sub>2</sub><sup>20</sup>** 8.64. 1,2,3-Benzotriazole crystallises from toluene, CHCl<sub>3</sub>, Me<sub>2</sub>NCHO or a saturated aqueous solution, and is dried at room temperature or in a vacuum oven at 65°. Losses are less if the material is distilled in a vacuum. **CAUTION: may EXPLODE during distillation; necessary precautions must be taken.** [Damschroder & Peterson *Org Synth Coll Vol III* 106 1955, *Beilstein* **26** III/IV 93.]

**O-Benzotriazol-1-yl-N,N,N',N'-tetramethyluronium hexafluorophosphate (HBTU)** [94790-37-1] **M** 379.2, **m** 200° (dec), 250°, 254°(dec). Wash the salt with H<sub>2</sub>O (3x), CH<sub>2</sub>Cl<sub>2</sub> (3x), dry and recrystallise it from MeCN. Dry it in a vacuum and store it cold in the dark [Dourtoglou et al. *Tetrahedron Lett* 1269 1978, NMR: Dourtoglou and Gross *Synthesis* 572 1984].

**Benzoxazolinone (2-hydroxybenzoxazole)** [59-49-4] **M** 135.1, **m** 137-139°, 142-143°(corrected), **b** 121-213°/17mm, 335-337°/760mm. Benzoxazolinone is purified by recrystallisation from aqueous Me<sub>2</sub>CO followed by distillation at atmospheric pressure, then in a vacuum. The *methyl mercury salt* recrystallises from aqueous EtOH and has **m** 156-158°. [Bywater et al. *J Am Chem Soc* **67** 905 1945, *Beilstein* **27** III/IV 2677.]

**S-(+) and R(-) 1-Benzoyl-2-tert-butyl-3-methyl-4-imidazolinone** [R- 101055-57-6] [S-101055-56-5] **M** 260.3, **m** 142-143°, 145.6-146.6°, 145-147°,  $[\alpha]_{546}^{20}$  (+) and (-) 155°,  $[\alpha]_D^{20}$  (+) and (-) 133° (c 1, CHCl<sub>3</sub>). Recrystallise these chiral imidazolinones from boiling EtOH (solubility is 1.43g/ml) or better by dissolving in CH<sub>2</sub>Cl<sub>2</sub> and adding pentane, filtering and drying for at least 12 hours at 60°/0.1mm and sublime them at 135°/0.01mm. They have also been purified by flash column chromatography through Merck silica gel at 0.04-0.063mm and using Et<sub>2</sub>O/petroleum ether/MeOH (60:35:5) as eluent. They are then recrystallised from EtOH/petroleum ether. [IR, NMR: Seebach et al. *Helv Chim Acta* **70** 237 1987, Fitzi & Seebach *Angew Chem, Int Ed Engl* **25** 345 1986.] The *racemate* is purified in a similar manner and has **m** 104-105° [NMR: Seebach et al. *Helv Chim Acta* **68** 949 1985].

**2-Benzoylpyridine** [91-02-1] **M** 183.2, **m** 41-43°, 48-50°, 72°/0.02mm, 104-105°/0.01mm,  $n_D^{24}$  1.6032, **pK<sub>Est</sub>** ~2.4. Dissolve 2-benzoylpyridine in Et<sub>2</sub>O, shake it with aqueous NaHCO<sub>3</sub>, H<sub>2</sub>O, dry it over MgSO<sub>4</sub> and evaporate. The residue solidifies on cooling. The solid can be recrystallised from petroleum ether. Its *hydrochloride* crystallises from Me<sub>2</sub>CO, **m** 126-127°, and the *2,4-dinitrophenylhydrazone* has **m** 193-195°. It distils at high vacuum. [Kinkerton & Thames *J Organomet Chem* **24** 623 1970, *Beilstein* **21/8** V 566.]

**N<sup>6</sup>-Benzyladenine** [1214-39-7] **M** 225.3, **m** 231-232°, 232.5°(dec), **pK<sub>Est(1)</sub>** ~ 4.2, **pK<sub>Est(2)</sub>** ~ 10.1. It is purified by recrystallisation from aqueous EtOH. It has  $\lambda_{max}$  at 207 and 270nm (H<sub>2</sub>O), 268 nm (pH 6), 274nm (0.1 N HCl) and 275nm (0.1 N NaOH). [Daly *J Org Chem* **21** 1553 1956, Bullock et al. *J Am Chem Soc* **78** 3693 1956, *Beilstein* **26** III/IV 3575.]

**1-Benzyl-1-aza-12-crown-4 (10-benzyl-1,4,7-trioxa-10-azacyclododecane)** [84227-47-4] **M** 265.4, **122-125°/0.03mm**, **140-143°/0.05mm**,  $d_4^{20}$  1.09,  $n_D^{20}$  1.52, **pK<sub>Est</sub>** ~ 7.7. Dissolve it in CH<sub>2</sub>Cl<sub>2</sub> or CCl<sub>4</sub> (1g in 30ml), wash it with H<sub>2</sub>O (30ml), brine (30ml), H<sub>2</sub>O (30 ml) again, dry (MgSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>), and evaporate. The residue in CH<sub>2</sub>Cl<sub>2</sub> is chromatographed through Al<sub>2</sub>O<sub>3</sub> (eluting with 10% EtOAc in hexane); evaporate, collect the correct fractions and distil (Kügelrohr) them. Log K<sub>Na</sub> in dry MeOH at 25° for Na<sup>+</sup> complex is 2.08. [White et al. *Tetrahedron Lett* **26** 151 1985, Arnold et al. *J Org Chem* **53** 5652 1988.]

**2-Benzyl-1,3-dioxolane** [101-49-5] **M 164.2, b 98-99°/1mm, 110°/5mm, 137-138°/34mm, 240-242°/atm,  $d_4^{20}$  1.087,  $n_D^{20}$  1.532.** Dissolve 2-benzyl-1,3-dioxolane in  $\text{CH}_2\text{Cl}_2$ , wash well with 1M NaOH, dry over  $\text{K}_2\text{CO}_3$ , filter, evaporate and distil it through a short path still (Kügelrohr). It has also been purified by preparative gas chromatography. [Raber & Guida *Synthesis* 808 1974, Lloyd & Luberoff *J Org Chem* 34 3949 1969, Beilstein 19 III/IV 220.]

**S-(+)- and R-(-)- Benzyl glycidyl ether (1-benzyloxyoxirane)** [S:14618-80-5] [R:16495-13-9] **M 164.2, b 68°/10<sup>-4</sup> mm, 105°/0.4mm,  $d_4^{20}$  1.072,  $n_D^{20}$  1.517,  $[\alpha]_{546}^{20}$  (+) and (-) 5.5°,  $[\alpha]_D^{20}$  (+) and (-) 5.1° (c 5, toluene),  $[\alpha]_D^{20}$  (+) and (-) 1.79° (c 5.02,  $\text{CHCl}_3$ ),  $[\alpha]_D^{21}$  (+) and (-) 15.3° (neat).** This ether in EtOAc is dried ( $\text{Na}_2\text{SO}_4$ ), then purified by flash chromatography using petroleum ether/EtOAc (5:1) as eluent. The ether distils through a short path distillation apparatus (Kügelrohr) as a colourless liquid. *Alternatively*, dissolve it in  $\text{CHCl}_3$ , wash it with  $\text{H}_2\text{O}$ , dry ( $\text{Na}_2\text{SO}_4$ ), evaporate and purify by silica gel chromatography. [Anisuzzamen & Owen *J Chem Soc* 1021 1967, Takano et al. *Heterocycles* 16 381 1981, Lipshutz et al. *Org Synth* 69 82 1990, Takano et al. *Synthesis* 539 1989, Honda et al. *Chem Pharm Bull Jpn* 39 1385 1991, Beilstein 12 IV 2277.]

**3-Benzyl-5-(2-hydroxyethyl)-4-methylthiazolinium chloride** [4568-71-2] **M 269.8, m 142-144°, 145-147°.** Purify the chloride by recrystallisation from EtOH or  $\text{H}_2\text{O}$ . If placed in a bath at 125° and heated at 2°/minute, the melting point is 140.5-141.4°. [Livermore & Sealock *J Biol Chem* 167 699 1947, Maier & Metzler *J Am Chem Soc* 79 4386 1957, Beilstein 27 III/IV 1758.]

**5-Benzyloxyindole** [1215-59-4] **M 223.3, m 96-97°; 100-103°, 104-106°,  $pK^{25}$  <0.** It is recrystallised from  $^*C_6H_6$ /petroleum ether or petroleum ether. The *picrate* forms red crystals from  $^*C_6H_6$  and has **m 142-143°.** [Burton & Leong *Chem Ind (London)* 1035 1953, Ek & Witkop *J Am Chem Soc* 76 5579 1954, fluorescence: Bridges & Williams *Biochem J* 107 225 1968, Beilstein 27 III/IV 1758.]

**1-Benzyl-4-piperidone** [3612-20-2] **M 189.3, b 107-108°/0.2mm, 114-116°/0.3mm, 143-146°/5mm, 157-158°/11mm,  $d_4^{20}$  1.059,  $n_D^{20}$  1.538.** If the physical properties show contamination, then dissolve it in the minimum volume of  $\text{H}_2\text{O}$ , made strongly alkaline with aqueous KOH, extract it with toluene several times, dry the extract with  $\text{K}_2\text{CO}_3$ , filter, evaporate and distil the residue at high vacuum using a bath temperature of 160-190°, and redistil it. [Brookes & Walker *J Chem Soc* 3173 1957, Bolyard *J Am Chem Soc* 52 1030 1930.] The *hydrochloride* has **m 159-161°** (from  $\text{Me}_2\text{CO}/\text{Et}_2\text{O}$ ), and the *picrate* has **m 174-182°** (from  $\text{Me}_2\text{CO}/\text{Et}_2\text{O}$ ). [Grob & Brenneisen *Helv Chim Acta* 41 1184 1958, Beilstein 21/6 V 424.]

**2-Benzylpyridine** [101-82-6] **M 169.2, b 98.5°/4mm,  $d_4^{20}$  1.054,  $n_D^{26}$  1.5771,  $pK^{25}$  5.13.** Dry it with NaOH for several days, then distil it from CaO under reduced pressure, and redistil the middle fraction. [Beilstein 20/7 V 556.]

**4-Benzylpyridine** [2116-65-6] **M 169.2, b 110.0°/6mm,  $d_4^{20}$  1.065,  $n_D^{26}$  1.5814,  $pK^{25}$  5.59.** Dry it with NaOH for several days, then distil it from CaO under reduced pressure, and redistil the middle fraction. [Beilstein 20/7 V 561.]

**Berbamine** [478-61-5] **M 608.7, m 197-210°,  $[\alpha]_D^{20}$  +115° ( $\text{CHCl}_3$ ),  $pK^{20}$  7.33. (70% aqueous EtOH).** Crystallise berbamine from petroleum ether or  $^*C_6H_6$  (with  $1^*C_6H_6$  **m 129-134°**). [Bick *Aust J Chem* 9 111, 118 1956, Beilstein 27 II 891, 27 III/IV 8732.]

**Berberine** [2086-83-1] **M 336.4, m 145°, 147-148°,  $pK_1^{20}$  2.47,  $pK_2^{20}$  11.73 (pseudobase?).** Berberine crystallises from petroleum ether or ether as yellow needles or from  $\text{H}_2\text{O}$ . [Beilstein 27 II 567, 27 III/IV 6539.]

**Berberine chloride (2H<sub>2</sub>O) (Neutral Yellow 18)** [633-65-8 (anhydrous), 5956-60-5 (2H<sub>2</sub>O)] **M 407.9, m 204-206°(dec), CI 75160,  $pK^{20}$  2.47.** Berberine chloride crystallises from water to give the *dihydrate*. The anhydrous salt may be obtained by recrystallisation from EtOH/ $\text{Et}_2\text{O}$ , wash the crystals with  $\text{Et}_2\text{O}$  and dry them in a vacuum. The *iodide* has **m 250°(dec)** (from EtOH). [Perkin *J Chem Soc* 113 503 1918, Kametani et al. *J Chem Soc(C)* 2036 1969, Beilstein 27 I 515, 27 II 567.]

**Bicuculline** See entry in “Miscellaneous”, Chapter 7.

**Bilirubin** [635-65-4] **M 584.7, m >360°**,  $\epsilon_{450\text{nm}}$  **55,600 in CHCl<sub>3</sub>, pK<sub>Est</sub> ~3.0**. An acyclic tetrapyrrole bile pigment with impurities which can be eliminated by successive Soxhlet extraction with diethyl ether and MeOH. It crystallises from CHCl<sub>3</sub> as deep red-brown rhombs, plates or orange-red prisms from chlorobenzene (**m 330° dec**) and is dried to constant weight at 80° under vacuum. [Gray et al. *J Chem Soc* 2264, 2276 1961, *Beilstein* 26 III/IV 3268.]

**Biliverdine** [114-25-0] **M 582.6, m >300°, pK<sup>25</sup> 3.0**. This is the precursor of bilirubin (above) and forms dark green plates or prisms, with a violet reflection, from MeOH. The *dimethyl ester*, when crystallised from MeOH has **m 215°(209°)**, and when crystallised from CHCl<sub>3</sub>/petroleum ether gives blue-green crystals with **m 202°**. [Gray et al. *J Chem Soc* 2264 1961, Sheldrick *J Chem Soc, Perkin Trans 2* 1457 1976, *Beilstein* 26 III/V 3272.]

**2-(4-Biphenyl)-5-phenyl-1,3,4-oxadiazole (BPD)** [852-38-0] **M 298.4, m 166-167°, 167-170°**. BPD is recrystallised from toluene. It is a good scintillation material [Brown et al. *Discussion Faraday Soc* 27 43 1959]. [*Beilstein* 27 III/IV 7283.]

**2,2'-Bipyridyl** [366-18-7] **M 156.2, m 70.5°, b 273°, pK<sub>1</sub><sup>25</sup> -0.52, pK<sub>2</sub><sup>25</sup> 4.44**. 2,2'-Bipyridyl crystallises from hexane, or EtOH, or (after charcoal treatment of a CHCl<sub>3</sub> solution) from petroleum ether. Also, it precipitates from a concentrated solution in EtOH by addition of H<sub>2</sub>O. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. It can be further purified by chromatography on Al<sub>2</sub>O<sub>3</sub> or by sublimation. Its UV (EtOH) has  $\lambda_{\text{max}}$  at 280nm (log $\epsilon$  4.13). [Airoldi et al. *J Chem Soc, Dalton Trans* 1913 1986, *Beilstein* 23 H 199, 23/8 V 16.]

**4,4'-Bipyridyl** [553-26-4] **M 156.2, m 73°(hydrate) [123333-55-1], 114° (171-171°)(anhydrous), b 305°/760mm, 293°/743mm, pK<sub>1</sub><sup>20</sup> 3.17, pK<sub>2</sub><sup>20</sup> 4.82**. It crystallises from water, \*benzene/petroleum ether, ethyl acetate and sublimes *in vacuo* at 70°. Also purify it by dissolving in 0.1M H<sub>2</sub>SO<sub>4</sub> and twice precipitating by addition of 1M NaOH to pH 8. Then recrystallise it from EtOH. For the *dihydrochloride* see Viologen below. [Collman et al. *J Am Chem Soc* 109 4606 1987, *Beilstein* 23 H 800, 23 III/IV 1371, 23/8 V 28.]

**2,2'-Biquinolin-4,4'-dicarboxylic (2,2'-bicinchoninic) acid** [1245-13-2] **M 344.3, m 367°, pK<sub>Est(1)</sub> ~1.5, pK<sub>Est(2)</sub> ~4.0**. Dissolve the acid in dilute NaOH and precipitate it with acetic acid, filter, wash swell with H<sub>2</sub>O and dry it at 100° in a vacuum oven. Attempts to form a picrate failed. The *methyl ester* (SOCl<sub>2</sub>-MeOH) has **m 165.6-166°**. [Lesesne & Henze *J Am Chem Soc* 64 1897 1942, Brown et al. *J Am Chem Soc* 68 2705 1946, *Beilstein* 25 III/IV 1148.] For the di-K salt see entry in “Metal-organic Compounds” in Chapter 5.

**2,2'-Biquinolyl ( $\alpha,\alpha'$ -diquinolyl)** [119-91-5] **M 256.3, m 196°, pK<sub>Est</sub> ~4.2**. Decolourise 2,2'-biquinolyl in CHCl<sub>3</sub> solution (charcoal), then crystallise it to constant melting point from EtOH or petroleum ether [Cumper et al. *J Chem Soc* 1188 1962]. [*Beilstein* 23/10 V 8.]

**2,5-Bis(4-aminophenyl)-1,3,4-oxadiazole (BAO)** [2425-95-8] **M 252.3, m 252-255°, 254-255°**. BAO is recrystallised from EtOH using charcoal and under N<sub>2</sub> to avoid oxidation. It is a fluorescent stain for DNA [Yataghanas et al. *Exptl Cell Res* 56 59 1969]. [*Beilstein* 27 III/IV 8158.]

**2,5-Bis(2-benzothiazolyl)hydroquinone** [33450-09-8] **M 440.3, m dec >200°**. Purify the hydroquinone by repeated crystallisation from dimethylformamide followed by sublimation in vacuum [Erusting et al. *J Phys Chem* 91 1404 1987].

**2,5-Bis(4-biphenyl)-1,3,4-oxadiazole (BBOD)** [2043-06-3] **M 374.5, m 229-230°, 235-238°**. BBOD is recrystallised from heptane or toluene. It is a good scintillant. [Hayes et al. *J Am Chem Soc* 77 1850 1955.]

**3,3-Bis(chloromethyl)oxacyclobutane (3,3-bis-[chloromethyl]oxetane)** [78-71-7] **M 155.0, m 18.9°, b 65°/5mm, 80°/10mm, 103°/30mm, 198°/atm, d<sub>4</sub><sup>20</sup> 1.290, n<sub>D</sub><sup>20</sup> 1.486**. Shake it with aqueous NaHCO<sub>3</sub> or FeSO<sub>4</sub> to remove peroxides, separate, dry with anhydrous Na<sub>2</sub>SO<sub>4</sub>, then distil it under reduced pressure from a little CaH<sub>2</sub> [Dainton et al. *Trans Faraday Soc* 56 1784 1960, Farthing *J Chem Soc* 3648 1955]. The 3,3-bis-

(*phenoxyethyl*) derivative is described below. [Beilstein 17 III/IV 68.] **Lachrymatory.**

***N,N'*-Bis(nicotinic acid) hydrazide** [840-78-8] **M 227-228°**, **m dec 200°**, **pK<sub>Est</sub> ~3.3**. The hydrazide crystallises from water; it is also soluble in hot EtOH but insoluble in \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or CHCl<sub>3</sub>. [Graf *J Prakt Chem* [2] 138 290 1933, Beilstein 22 III/IV 455.]

**3,3-Bis(phenoxyethyl)oxacyclobutane (3,3-bis-[phenoxyethyl]oxetane)** [1224-69-7] **M 270.3**, **m 67.5-68°**, **b 143°/0.05mm**. Distil it under high vacuum, then crystallise the solidified distillate from MeOH. [Farthing *J Chem Soc* 3648 1955, Beilstein 17 III/IV 2010.]

**Blue Tetrazolium (tetrazolium blue chloride)** [1871-22-3] **M 727.7**, **m 254-255°(dec)**. Crystallise the chloride from 95% EtOH/anhydrous diethyl ether to constant absorbance at 254nm. [Beilstein 26 III/IV 1789.]

**1-*N*-Boc-piperidine (1-*N*-*tert*-butoxycarbonylpiperidine, 1-piperidincarboxylic acid 1,1-dimethylethyl ester)** [75844-69-8] **M 185.3**, **b 79°/0.05mm**, **d<sub>4</sub><sup>25</sup> 0.964**, **n<sub>D</sub><sup>20</sup> 1.454**. It is purified by bulb-to-bulb distillation at 65°/1mm. If it is discoloured then dissolve it in Et<sub>2</sub>O, wash it with brine, separate, dry the Et<sub>2</sub>O layer over K<sub>2</sub>CO<sub>3</sub>, filter, evaporate, and distil in a vacuum. Its IR has  $\nu_{\max}$  at 1695 (CO) cm<sup>-1</sup> (film). [IR, NMR, MS: Dieter & Li *J Org Chem* 62 7726 1997, Beak & Lee *J Org Chem* 58 1109 1993].

**1-*N*-Boc-2-piperidone (1-*N*-*tert*-butoxycarbonyl-2-piperidone)** [85908-96-9] **M 199.3**, **m 30-33°**, **32-34°**, **33-35°**, **36°**, **b 110°/0.1mm**. This useful synthetic starting material is prepared by using the same principles as for the 4-oxo isomer below. Thus to a solution of  $\gamma$ -valerolactam (2-piperidone, 6.0g, 60.5mmol, [675-20-7]), DMAP (1.85g, 15.1mmol, 0.2 equiv, [1122-58-3]) and Boc<sub>2</sub>O (26.6g, 121.0mmol, 2.0 equiv, [24074-26-8]) in CH<sub>2</sub>Cl<sub>2</sub> (70ml) at ~25° is added Et<sub>3</sub>N (17.8g, 127.7mmol, 2.1 equiv) and is stirred for 32 hours, then quenched with 1.2N HCl (10ml). The aqueous layer is extracted with CH<sub>2</sub>Cl<sub>2</sub>, the combined CH<sub>2</sub>Cl<sub>2</sub> solutions are washed with saturated aqueous NaHCO<sub>3</sub> (20ml), brine (30ml), dried (MgSO<sub>4</sub>), filtered and evaporated *in vacuo*. The crude product is purified by flash chromatography through Kieselgel 60 (230-240 mesh with cyclohexane/EtOAc 8:2 elution) to give *1-N-Boc-2-piperidone* (9.71g, 48.7mmol, 81%) as a colourless low melting solid **m 36°**. It has R<sub>F</sub> 0.40 (TLC on Kieselgel 60F<sub>254</sub>, cyclohexane/EtOAc 7:3), and its IR (KBr) has  $\nu_{\max}$  at 1713, 1302, 1249, 1159 and 1138 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, with residual CHCl<sub>3</sub> as internal standard at  $\delta$  7.27) has  $\delta$  at 3.66 (m, 2H, H-6,6), 2.50 (m, 2H, H-3,3), 1.90-1.74 (4H, H-3,3,4,4) and 1.53 (s, 9H, *t*-Bu); the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>, with residual CDCl<sub>3</sub> as internal standard at  $\delta$  77.1) has  $\delta$  at 171.1 (s), 152.4 (s), 82.5 (s), 46.0 (t), 34.6 (t), 27.7 (q), 22.5 (t) and 20.2 (t); and the EI MS has *m/z* (relative intensity) at 199 (M + 0.3), 144 (52), 126 (21), 100 (38), 99 (31), 98 (31), 82 (24), 57 (100), 56 (38) and 55 (30). [Cossy et al. *New J Chem* 27 475 2003.] This synthesis has also been carried out in MeCN/DMAP at ~25°/24 hours (74% yield) without further base [Moody & Taylor *J Chem Soc, Perkin Trans 1* 721 1989].

*1-N-Boc-2-piperidone* is hydrolysed at ambient temperature by LiOH (30 minutes, 90%), or undergoes methanolysis by MeONa/MeOH (15 minutes, 94%) to provide  $\gamma$ -BocNH(CH<sub>2</sub>)<sub>4</sub>CO<sub>2</sub>H (after acidification) or  $\gamma$ -BocNH(CH<sub>2</sub>)<sub>4</sub>CO<sub>2</sub>Me without losing the Boc protecting group, respectively [Flynn et al. *J Org Chem* 54 2424 1989]

**1-*N*-Boc-3-piperidone (1-*N*-*tert*-butoxycarbonyl-3-piperidone)** [98977-36-7] **M 199.3**, **m 35-40°**, **b 104-105°/0.4mm**. 3-Piperidone used to prepare the Boc derivative was obtained by hydrogenolysis of *N*-benzyl-3-piperidone hydrochloride hydrate (4.2g, 18.6mmol, [50606-58-1]) catalysed by 10% Pd/C (0.8g) and H<sub>2</sub> at 55 psi while stirring for 16 hours in degassed MeOH (200ml). This was filtered (Celite) and evaporated *in vacuo*. The crude oily 3-piperidone was dissolved in THF (200ml), treated with Boc<sub>2</sub>O (5.27g, 24.1mmol) and saturated aqueous Na<sub>2</sub>CO<sub>3</sub> (50ml), stirred for 4 hours and evaporated *in vacuo*. The white solid was partitioned between EtOAc and 1N HCl, the organic layer was collected, washed with 1N NaOH and brine, dried (MgSO<sub>4</sub>), filtered, and evaporated *in vacuo* to give an oil which was purified by flash chromatography (silica gel, hexane/EtOAc 3:1) to give *1-N-Boc-3-piperidone* (2.93g, ~100%) as a colourless oil. [Lucca et al. *J Med Chem* 48 2195 2005.] Its <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 3.99 (s, 2H, 2,2-H), 3.58 (t, *J* = 6.3 Hz, 2H, 6,6-H), 2.46 (t, *J* = 6.3 Hz, 2H 4,4-H), 1.97 (t, *J* = 6.3 Hz, 2H, 5,5-H), and 1.45 (s, 9H, *t*-Bu); and the HRMS has

$m/z$  200.1285 and calculated for  $C_{10}H_{18}NO_3$  (M + H) was 200.1287. *Alternatively*, and almost similar preparation, was carried out except that saturated aqueous  $Na_2CO_3$  was replaced by  $Et_3N$  (3 equiv) to give an 81% yield of *1-N-Boc-3-piperidone* which was distilled at high vacuum undecomposed. [Brehm et al. *J Med Chem* **29** 224 1986.]

In a completely different synthesis, involving catalysed intramolecular cyclisation, (*5-oxo-6-dimethylsulfoxonium-hexyl*)-carbamic acid *tert-butyl ester* (250mg, 0.901mmol) in  $CH_2Cl_2$  (8.0ml, purged with  $N_2$  for 30 minutes) was added *via* a syringe pump over 18 hours to a solution of  $[Ir(COD)Cl]_2$  (6.05mg, 9.01  $\mu$ mol, 0.01 equiv, [12112-67-3]) in degassed  $CH_2Cl_2$  (4.0ml) at 70°. The mixture was then evaporated *in vacuo*, and the residue was purified through a silica gel column and eluting with hexanes/EtOAc (3:2) to afford *1-N-Boc-3-piperidone* (147mg, 82%) identical with the above. [Mangion et al. *Org Lett* **11** 3566 2009.]

**1-N-Boc-4-piperidone (1-N-*tert*-butoxycarbonyl-4-piperidone)** [98977-36-7] **M 199.3, m 68-69°, 70-72°, 72°, 73-75°, 74-75°, 74.4-75.2°**. It is prepared by dissolving 4-piperidone (1g, 6.5mmol [41661-47-6]) and  $EtN(iso-Pr)_2$  (2.8ml, 16mmol, DIPEA [7087-68-5]) in dioxane/ $H_2O$  (4:1, 10ml), adding  $Boc_2O$  (2.2g, 9.8mmol, [24074-26-8]) slowly with stirring at ~25° for 24 hours. The solvent is removed *in vacuo*, the residue is dissolved in  $CH_2Cl_2$ , washed with  $H_2O$  (3 x), dried ( $MgSO_4$ ), filtered and evaporated *in vacuo*. The residue is triturated with  $Et_2O$ , the white residue is collected by filtration and recrystallised from hexane to give the *Boc-piperidone* (75%), **m 73.5°**. It also crystallises from petroleum ether, toluene or EtOAc and can be dried in a vacuum desiccator over shredded paraffin wax. It has  $R_F$  0.74 (TLC on Kieselgel 60F<sub>254</sub>, hexane/EtOAc 4:6), and its IR (KBr) has  $\nu_{max}$  at 1715 and 1680  $cm^{-1}$ . Its  $^1H$  NMR (300MHz,  $CDCl_3$ ) has  $\delta$  at 3.72 (t,  $J = 6.3$  Hz, 4H, 3,5-H), 2.44 (t,  $J = 6.3$  Hz, 4H, 2,6-H) and 1.50 (s, 9H, *t*-Bu); and the  $^1H$  NMR (300MHz,  $DMSO-d_6$ ) has  $\delta$  at 3.6 (t, 4H, 3,5-H), 2.4 (t, 4H, 2,6-H) and 1.4 (s, 9H, *t*-Bu). It has been used extensively to introduce a piperidine ring for making drugs because the 4-oxo group can be functionalised in various ways and deprotection of the nitrogen allows bonding to this atom. [Houssin et al. *J Med Chem* **45** 533 2002, Ellis et al. *J Med Chem* **51** 2170 2002.]

**Brazilin (6a*S*-*cis*-7,11b-dihydrobenzo[*b*]indeno[1,2-*d*]pyran-3,6a,9,10(6*H*)-tetraol)** [474-07-7] **M 269.3, m 130°(dec), 250°,  $pK_{Est(1)} \sim 9.3$ ,  $pK_{Est(2)} \sim 10.0$ ,  $pK_{Est(3)} \sim 12.5$  (all phenolic), CI 75280**. Brazilin crystallises from EtOH as yellow crystals which become orange when exposed to light and air, and is yellow in dilute acid but crimson in dilute alkali. When crystallised from  $H_2O$ , it has **m 247-248°**. It forms coloured metal salts and is oxidised in air to *Brazilein*, the quinonoid form. The ( $\pm$ )-*form has been resolved, and the (+)-enantiomer* has  $[\alpha]_D^{20} +121^\circ$  (c 1, MeOH). [Craig et al. *J Org Chem* **30** 1573 1965, Morsingh & Robinson *Tetrahedron* **26** 281 1970, Beilstein **17** H 194, **17** II 244, **17** III/IV 2711.]

**Brilliant Cresyl Blue (N*N*-diethyl-3-imino-8-methyl-3*H*-phenoxazin-7-amine hydrochloride)** [4712-70-3] **M 332.8,  $pK^{25} 3.2$** . Crystallise the dye from petroleum ether. It has  $\lambda_{max}$  at 625nm ( $H_2O$ ) and 622nm (95% aqueous EtOH). [Beilstein **27** H 400, **27** II 454, **27** III/IV 5160.]

**5-Bromocytosine** [2240-25-7] **M 190.0, m 245-255°(dec), 250°(dec),  $pK_1^{25} 3.04$ ,  $pK_2^{25} 10.33$** . 5-Bromocytosine is recrystallised from  $H_2O$  or 50% aqueous EtOH. *Alternatively*, dissolve *ca* 3g in conc HCl (10ml) and evaporate to dryness. Dissolve the residual hydrochloride in the minimum volume of warm  $H_2O$  and make faintly alkaline with aqueous  $NH_3$ . Collect the crystals and dry them in a vacuum at 100°. [Hilbder & Jensen *J Am Chem Soc* **56** 134 1934, Beilstein **25** III/IV 3689.]

**2-(2-Bromoethyl)-1,3-dioxane** [33884-43-4] **M 195.1, b 67-70°/2.8mm, 71-72°/4mm, 95°/15mm,  $d_4^{20} 1.44$ ,  $n_D^{20} 1.4219$** . Purify it by vacuum fractionation. Also dissolve it in  $Et_2O$ , wash with aqueous  $NaHCO_3$ , dry the extract ( $Na_2SO_4$ ), filter and fractionate at high vacuum. Its  $^1HNMR$  in  $CCl_4$  has  $\delta$  at 1.3 (m, 1H), 2.1 (m, 3H), 3.36 (t, 2H), 3.90 (m, 4H) and 4.57 (t, H). [Stowell *J Org Chem* **41** 560 1976, NMR, MS: Schwarz & Franke *Tetrahedron* **35** 1969 1979, Kriesat & Gisvold *J Pharm Sci* **60** 1250 1971, Beilstein **19/1** V 69.]

**2-(2-Bromoethyl)-1,3-dioxolane** [18742-02-4] **M 181.1, b 68-80°/8mm, 68-73°/10mm, 78-80°/20mm,  $d_4^{20}$**

**1.510, n<sub>D</sub><sup>20</sup> 1.479.** Dissolve it in pentane, wash with 5% aqueous NaHCO<sub>3</sub>, dry (Na<sub>2</sub>SO<sub>4</sub>), and evaporate. Distil the residue. [NMR: Büchi & Wüest *J Org Chem* **34** 1122 1969, Kriesat & Gisvold *J Pharm Sci* **60** 1250 1971, *Beilstein* **19/1** V 69.]

**6-Bromo-5-fluoro-1H-indole** [259860-08-7] **M 214.0, m 83-84°, pK<sub>Est(1)</sub> ~ -4.05 (basic), pK<sub>Est(2)</sub> ~15.9 (acidic).** There are several preparations of this indole. In a patent (to Vernallis Research Ltd, US6380238, 2002) *N,N*-dimethylformamide dimethylacetal (8.5ml, 60mmol) was added to a stirred solution of 4-bromo-5-fluoro-2-nitrotoluene (11.8g, 50mmol) in DMF (30ml) in one portion under argon at ~25°, then heated at 120° for 16 hours and evaporated *in vacuo*. The residual oil crystallised from MeOH/CH<sub>2</sub>Cl<sub>2</sub> (4:1) to give a purple solid (4.5g) which was dissolved in MeOH/THF (1:1, 30ml), mixed with Raney Ni (1g), cooled to 0° and N<sub>2</sub>H<sub>4</sub>·H<sub>2</sub>O (0.8ml, 16mmol) was added all at once. After stirring for 90 minutes, more N<sub>2</sub>H<sub>4</sub>·H<sub>2</sub>O (0.8ml) was added and stirred at 0° for 30 minutes. The mixture was filtered through Celite and the filter cakes was washed with THF. The combined filtrates were evaporated *in vacuo*, and the residue was purified by chromatography [SiO<sub>2</sub> with heptane/CH<sub>2</sub>Cl<sub>2</sub> (4:1) as eluent] to give the *indole* (1.7g, 16%) as an off-white solid. In a second patent [to M.P. Dillon et al. US2004/224973, 2004, see also Batcho & Leimgruber *Org Synth* **63** 214 1985] crude [2-(4-bromo-5-fluoro-2-nitrophenyl)vinyl]-dimethylamine and Raney Ni in THF gave after flash chromatography through a silica gel column (eluting with 10% EtOAc in hexane) which gave the indole as a light green solid (61%). However, when the same enamine (1.89g, 4.38mmol) in EtOH (35ml) was treated with a mixture of Fe (4.42g, 79mmol) in AcOH (35ml) and stirred at 90° overnight, then filtered, evaporated and the residue purified by chromatography as above, pure *6-bromo-5-fluoro-1H-indole* (0.83g, 89%) was obtained as a yellow solid [Patent to the Board of Trustees of the University of Illinois; WO2008/77138 2008]. It had IR (nujol) with  $\nu_{\max}$  at 3395, 2925, 2855, 1570, 1469, 1451, 1408, 1314, 1145, 1105, 865, 763 and 505 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) had  $\delta$  at 7.85 (br s, 1H), 7.55 (dd, 1H, *J* = 5.6 and 0.9 Hz), 7.34 (d, 1H, *J* = 9.0 Hz), 7.23 (t, 1H, *J* = 2.8 Hz) and 6.45-6.51 (m, 1H).

*6-Bromo-5-fluoro-1-methylindole* was obtained as a colourless solid in 99% yield by converting the indole to its *N-sodio-* derivative with NaH (1.5mol) in DMF at 0°/30 minutes, followed by MeI (1.2mol) for 1 hour at ~25°, then purified by column chromatography as above. Its <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) had  $\delta$  at 3.66 (s, 3H, *N-Me*), 6.43 (d, 1H, *J* = 2.7 Hz), 7.05 (d, 1H, *J* = 2.7 Hz), 7.34 (d, 1H, *J* = 9.3 Hz) and 7.44 (d, 1H, *J* = 5.4 Hz) [Patent to the Board of Trustees of the University of Illinois; WO2008/77138 2008].

**3-Bromofuran** [22037-28-1] **M 147.0, b 38.5°/40mm, 50°/110mm, 102.5-103°/atm, d<sub>4</sub><sup>20</sup> 1.661, n<sub>D</sub><sup>20</sup> 1.4970.** Purify 3-bromofuran by two steam distillations and dry it over fresh CaO. It can be dried over Na metal (no obvious reaction) and fractionated. It is not very soluble in H<sub>2</sub>O but is soluble in organic solvents. When freshly distilled, it is a clear oil, but darkens on standing and eventually resinifies. It can be stored for long periods by covering the oil with an alkaline solution of hydroquinone and is redistilled when required. It forms a characteristic *maleic anhydride adduct*, **m 131.5-132°**. [Shepard et al. *J Am Chem Soc* **52** 2083 1930, Huhes & Johnson *J Am Chem Soc* **53** 737 1931, adduct: van Campen & Johnson *J Am Chem Soc* **55** 430 1933, *Beilstein* **17/1** V 295.]

**5-Bromoindole** [10075-50-0] **M 196.1, m 90.5-91°, 90-92°, pK<sup>25</sup> 16.13 (NH).** Purify it by steam distillation from a faintly alkaline solution. Cool the aqueous distillate, collect the solid, dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> and recrystallise it from aqueous EtOH (35% EtOH) or petroleum ether/Et<sub>2</sub>O. Its UV in MeOH has  $\lambda_{\max}$  at 279, 287 and 296nm (log $\epsilon$  3.70, 3.69 and 3.53). The *picrate* has **m 137-138°(dec)** (from Et<sub>2</sub>O/petroleum ether). [UV: Thesing et al. *Chem Ber* **95** 2205 1962, UV and NMR: Lallemand & Bernath *Bull Soc Chim Fr* 4091 1970, *Beilstein* **20/7** V 36.]

**5-Bromoisatin** [87-48-9] **M 226.0, m 245°(dec), 251-153°, 255-256°.** 5-Bromoisatin forms red prisms or needles from EtOH. The *N-acetate* crystallises as yellow prisms from \*C<sub>6</sub>H<sub>6</sub>, **m 170-172°**, and the *N-methyl* derivative forms orange-red needles from MeOH, **m 172-173°**. [Heller *Chem Ber* **53** 1545 1920, Buu-Hoi *Rec Trav Chim Pays Bas* **73** 197 1954, Baker et al. *Tetrahedron Lett* 215 1978, *Beilstein* **21** H 453, **21** III/IV 5009.]

**6-Bromoisatin** [6326-79-0] **M 226.0, m 270°, pK<sup>25</sup> 10.35.** 6-Bromoisatin recrystallises from AcOH (yellow needles). It is a plant growth substance. Its IR (CHCl<sub>3</sub>) has  $\nu_{\max}$  at 1320 and 3440cm<sup>-1</sup>. [Sadler *J Org Chem*

21 169 1956, *Beilstein* 21 III/IV 5012.]

**2-Bromo-3-methylindole (2-bromoskatole)** [1484-28-2] **M 210.1, m 102-104°**,  $\text{pK}_{\text{Est}} < 0$ . Purify 2-bromoskatole by chromatography on silica gel in  $\text{CHCl}_3$ /petroleum ether (1:2) followed by crystallisation from aqueous EtOH. [Phillips & Cohen *J Am Chem Soc* 108 2023 1986, cf *Beilstein* 20 III/IV 3205.]

**4-(Bromomethyl)-7-methoxycoumarin** [35231-44-8] **M 269.1, m 208-209°, 213-215°, 216-218°**. The coumarin is crystallised from boiling AcOH, the crystals are washed with AcOH, EtOH and dried in a vacuum; The  $^1\text{H NMR}$  (TFA) has  $\delta$  at 3.97s, 4.57s, 6.62s, 6.92-7.19m and 7.80d. [Secrist et al. *Biochem Biophys Res Commun* 45 1262 1971, Dünge *Anal Chem* 49 442 1977, *Beilstein* 18 III/IV 348.]

**5-Bromonicotinic acid** [20826-04-4] **M 202.0, m 178-182°, 189-190°,  $\text{pK}_{\text{Est}} \sim 4.4$ ,  $\text{pK}^{25} 4.02$  (50% aqueous EtOH)**. The acid is recrystallised from  $\text{H}_2\text{O}$  and then from EtOH using charcoal. The *amide* has **m 219-219.5°** (from aqueous EtOH), and the *methyl ester*, prepared by addition of ethereal diazomethane, can be purified by sublimation in a vacuum and has **m 98-99°**. The *acid chloride* also can be sublimed *in vacuo* and has **m 74-75°** and gives the *methyl ester* in MeOH. [Graf *J Prakt Chem* 138 244 1933, Bachman & Micucci *J Am Chem Soc* 70 2381 1948, Garcia et al. *J Am Chem Soc* 82 4430 1960, Mistic-Vokovic et al. *J Chem Soc* 34 1978, *Beilstein* 22/2 V 181.]

**2-Bromopyridine** [109-04-6] **M 158.0, b 49.0°/2.7mm,  $d_4^{20} 1.660$ ,  $n_D^{20} 1.5713$ ,  $\text{pK}^{25} 0.90$** . Dry 2-bromopyridine over KOH for several days, then distil it from CaO under reduced pressure, and taking the middle fraction. [*Beilstein* 20/5 V 422.]

**8-Bromotheophylline (bromo-1,3-dimethyl-2,6(1*H*,3*H*)-purinedione)** [10381-75-6] **M 259.1, m 309°, 315-320° (with browning and dec)**,  $\text{pK}_{\text{Est}(1)} \sim 5.5$ ,  $\text{pK}_{\text{Est}(2)} \sim 9.2$ . It is purified by dissolving in the minimum volume of dilute NaOH (charcoal), filtering and acidifying to pH *ca* 3.5-4. The solid that separates is collected, dried *in vacuo* at 100° and stored in a dark container. It has also been recrystallised from EtOH or AcOH. [Blitz & Beck *J Prakt Chem* [2] 118 158 1928, Fischer & Ach *Chem Ber* 28 3142 1895, *Beilstein* 26 H 476, 26 II 227, 26 III/IV 2447.]

**5-Bromothiazole** [3034-55-7] **M 164.0, b 62-63°/15mm,  $d_4^{20} 1.835$ ,  $n_D^{20} 1.5955$ ,  $n_D^{25} 1.5976$ ,  $\text{pK}_{\text{Est}} \sim 2.0$** . If bromothiazole is too coloured, then suspend it in dilute NaOH and steam distil it. Add NaCl to the aqueous distillate, extract it with  $\text{Et}_2\text{O}$ , dry it ( $\text{Na}_2\text{SO}_4$ ), evaporate and fractionate the residue in a vacuum. The  $\text{HgCl}_2$  salt crystallises from EtOH with **m 148° (dec)**. [Beyerman et al. *Rec Trav Chim Pays Bas* 73 330 1954, *Beilstein* 27 III/IV 962.]

**2,3-Dibromothiophene** [3140-93-0] **M 241.9, m -17.5°, b 89-91°/13mm, 212-213°/atm, 218.6-219.6°/atm,  $d_4^{25} 2.137$ ,  $n_D^{20} 1.632$** . Purify the dibromothiophene by fractional distillation, preferably in a vacuum. Low temperature crystallisation from a small volume of petroleum ether (use  $\text{Et}_2\text{O}/\text{CO}_2$  bath) removes the more soluble 2,4-dibromothiophene isomer. Nitration with  $\text{Ac}_2\text{O}/\text{HNO}_3$  at 50-55° yields 2,3-dibromo-5-nitrothiophene **m 75°** (from EtOH). [Steinkopf et al. *Justus Liebigs Ann Chem* 512 149-151 1934, Gronowitz et al. *Acta Chem Scand* 46 654 1992, NMR: Fujikara et al. *Bull Chem Soc Jpn* 32 201 1959, *Beilstein* 17/1 III/IV 247, 17/1 V 308.]

**Bufotenine hydrogen oxalate (3-[2-{dimethylamino}ethyl]-5-hydroxyindole hydrogen oxalate)** [2963-79-3] **M 294.3, m 96.5°,  $\text{pK}_1^{25} \sim 4.9$  (estimated),  $\text{pK}_2^{25} 9.8$  (5-OH),  $\text{pK}_3^{25} 11.2$  ( $\text{NMe}_2$ ),  $\text{pK}_4^{25} \sim 18.2$  (estimated acidic indole NH)**. The hydrogen oxalate crystallises from  $\text{Et}_2\text{O}$ , EtOH/ $\text{Et}_2\text{O}$  (gave *monohydrate*) or MeOH/ $\text{Et}_2\text{O}$ . [Wieland et al. *Justus Liebigs Ann Chem* 513 11 1934, *Beilstein* 22 III/IV 5672.]

**4-*tert*-Butylcalix[4]arene** [60705-62-6] **M 648.9, m >300°(dec), 380°(dec), 344-346°**. The calixarene recrystallises from  $\text{CHCl}_3$  in large solvated prisms (**m 380° dec**); it effloresces on drying in air. Its *tetra-acetate* crystallises from  $\text{Ac}_2\text{O}$  in colourless prisms **m 332-333°(dec)**. It crystallises from  $\text{CCl}_4$  or chlorobenzene /EtOH (**m >300°**) and the *tetra-acetate* crystallises from  $\text{CHCl}_3$ /EtOH **m >290°(dec)**. It also crystallises from toluene in white plates with toluene of crystallisation **m 344-346° (330-332°)**; the *tetra-acetate* crystallises with 1AcOH of crystallisation **m 383-386° (softening at 330-340°, also m 283-286°)**, but acetylation with

Ac<sub>2</sub>O/NaOAc gives the *triacetate* which recrystallises from AcOH with 1AcOH of crystallisation **m** 278-281°. 4-*tert*-Butylcalix[4]arene (100mg) is unchanged after boiling for 4 hours with 10N KOH (0.04ml) in xylene (4ml). [*Br J Pharmacol* **10** 73 1955, Kämmerer et al. *Monatsh Chem* **109** 767 1978, Gutsche et al. *J Am Chem Soc* **103** 3782 1981; see also Kluawer in *Calixarenes*, Vicens & Böhner eds Academic Press 1991, *Beilstein* **6** IV 7858.]

4-*tert*-Butylcalix[6]arene [78092-53-2] **M** 972.3, **m** >300°, **380-381°**. It is recrystallised from CHCl<sub>3</sub> or CHCl<sub>3</sub>/MeOH to give a white solid from the mother liquors of the calix[8]arene preparation. The *hexa-acetate* (Ac<sub>2</sub>O/H<sub>2</sub>SO<sub>4</sub>) crystallises from CHCl<sub>3</sub>/MeOH with **m** 360-362°(dec), and the (SiMe<sub>3</sub>)<sub>6</sub> derivative crystallises from CHCl<sub>3</sub>/MeOH with **m** 410-412°. Its stability in KOH-xylene is the same as for the 4-*tert*-butylcalix[4]arene. [Gutsche et al. *J Am Chem Soc* **103** 3782 1981. See also Kluawer in *Calixarenes*, Vicens & Böhner eds Academic Press 1991, *Beilstein* **6** IV 7858.]

4-*tert*-Butylcalix[8]arene [68971-82-4] **M** 1297.8, **m** 411-412°. The calixarene recrystallises from CHCl<sub>3</sub> in fine colourless, glistening needles. It melts sharply between 400-401° and 411-412° depending on the sample and is sensitive to traces of metal ions. On TLC with silica gel (250µm thick) and elution with CHCl<sub>3</sub>/hexane (3:4) it has R<sub>F</sub> 0.75. The *octa-acetate* is prepared from 8g in Ac<sub>2</sub>O (50ml) and 2 drops of conc H<sub>2</sub>SO<sub>4</sub> and refluxed for 2 hours. On cooling, a colourless precipitate separates and is recrystallised from Ac<sub>2</sub>O (1.2g 48%) with **m** 353-354°. The (SiMe<sub>3</sub>)<sub>8</sub> is prepared from 4-*tert*-butylcalix[8]arene (0.65g) in pyridine (4ml) with excess of hexamethyldisilazane (1ml) and trimethylchlorosilane (0.5ml) and refluxed under N<sub>2</sub> for 2 hours. Cool, evaporate the pyridine, triturate the gummy residue with MeOH. Chromatograph on silica gel using hexane/CH<sub>2</sub>Cl<sub>2</sub> gave 0.5g (61%) with one spot on TLC. Recrystallise it from hexane/Me<sub>2</sub>CO to give colourless needles **m** 358-360°. [Gutsche et al. *J Am Chem Soc* **103** 3782 1981, Gutsche & MuthuKrishnan *J Org Chem* **43** 4905 1978, Gutsche & MuthuKrishnan *J Org Chem* **44** 3962 1979, Andretti et al. *J Chem Soc, Chem Commun* 533 1981; see Kluawer in *Calixarenes*, Vicens & Böhner eds Academic Press 1991.]

8-*sec*-Butylmetrazole [25717-83-3] **M** 194.3, **m** 70°. Crystallise it from petroleum ether and dry it for 2 days under vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **26** II 213 for Metrazole.]

*N*-(*n*-Butyl)-5-nitro-2-furamide [14121-89-2] **M** 212.2, **m** 89-90°, **b** 190°/10mm. Distil the amide in a vacuum and recrystallise it twice from EtOH/water mixture or petroleum ether. [Gliman & Yale *J Am Chem Soc* **72** 3593 1950, *Beilstein* **18** III/IV 3995.]

Butyloxirane (1-hexene oxide) [1436-34-6] **M** 100.2, **b** 116-117°/atm, 116-119°/atm, **d**<sub>4</sub><sup>20</sup> 0.833, **n**<sub>D</sub><sup>20</sup> 1.44051. Purify it by fractional distillation through a 2ft helices-packed column at atmospheric pressure in a N<sub>2</sub> atmosphere. [Pasto & Cumbo *J Org Chem* **30** 1271 1965, Howarth et al. *J Chem Soc* 2433 1927, <sup>13</sup>C NMR Davies & Witham *J Chem Soc Perkin Trans 2* 861 1975, *Beilstein* **17/1** V 103.]

4-*tert*-Butylpyridine [3978-81-2] **M** 135.2, **f** -44.4°, **b** 194-197°/atm, 197°/765mm, **d**<sub>4</sub><sup>20</sup> 0.923, **n**<sub>D</sub><sup>20</sup> 1.495, **pK**<sup>25</sup> 5.82. Dry 4-*tert*-butylpyridine over solid KOH and purify it by fractional distillation through an efficient column under dry N<sub>2</sub>. Its *picrate* has **m** 153.9-154°, and the *hydrochloride* has **m** 151.7-154.8° (from Me<sub>2</sub>CO). [Brown & Murphey *J Am Chem Soc* **73** 3308 1951, IR: Arnett & Chawla *J Am Chem Soc* **100** 214 1978, Kyle et al. *J Chem Soc* 4454 1960, *Beilstein* **20/6** V 123.]

**Cacotheline** (2,3-dihydro-4-nitro-2,3-dioxo-9,10-*secostrychnidin-10-oic acid*) [561-20-6] **M** 508.4, **pK**<sub>Est(1)</sub> ~4.4 (CO<sub>2</sub>H), **pK**<sub>Est(2)</sub> ~10.2 (N). Cacotheline gives yellow crystals from H<sub>2</sub>O. It is then dried over H<sub>2</sub>SO<sub>4</sub> which gives the *dihydrate*, and in a vacuum over H<sub>2</sub>SO<sub>4</sub> at 105° it forms the *anhydrous* compound. The *hydrochloride* separates as the *hydrate* (on heating in vacuum at 80°) in orange-yellow prisms or plates, **m** 250°(dec) and forms a *resorcinol complex* which gives brown crystals from EtOH, **m** 325°; and a *hydroquinone complex* as dark red crystals from EtOH, **m** 319°. [Leuchs & Leuchs *Chem Ber* **43** 1042 1910, Teuber *Chem Ber* **86** 232, (UV: 242) 1953; complexes: Gallo *Gazz Chim Ital* **85** 1441 1955.] It is used in the titrimetric estimation of Sn<sup>2+</sup> ions [Szrvas & Lantos *Talanta* **10** 477 1963]. [*Beilstein* **27** III/IV 8014.]

**Caffeine (1,3,7-trimethylxanthine)** [58-08-2] **M 194.2, m 237°, pK<sub>1</sub><sup>40</sup> -0.10, pK<sub>2</sub><sup>55</sup> 1.22.** Caffeine crystallises from water or absolute EtOH. [Beilstein 26 III/IV 2338.]

**Cannabinol** [521-35-7] **M 310.4, m 76-77°, b 185°/0.05mm, 225°, pK<sub>Est</sub> ~10.5 (phenolic OH).** Cannabinol crystallises from petroleum ether and sublimes in a vacuum. [Meitzer et al. *Synthesis* 985 1981, *Beilstein* 17 II 151, 17 III/IV 1652.]

**ε-Caprolactam (azepan-2-one, aza-2-cycloheptanone, 2-oxohexamethyleneimine)** [105-60-2] **M 113.2, m 70°, 70.5-71.5°, 70-71°, 262.5°/760mm.** The lactam is distilled under reduced pressure, recrystallised from acetone or petroleum ether and redistilled. It can be purified by zone melting. It is very *hygroscopic* and discolours in contact with air unless small amounts (0.2g/L) of NaOH, Na<sub>2</sub>CO<sub>3</sub> or NaBO<sub>2</sub> are present. It has been crystallised from a mixture of petroleum ether (185ml of **b** 70°) and 2-methyl-2-propanol (30ml), from acetone, or petroleum ether. It is then distilled under reduced pressure and stored under nitrogen. [Pellegata et al. *Synthesis* 614 1978, *Beilstein* 21/6 V 444.]

**Caprylolactam (azanon-2-one, azacyclononan-2-one, 8-aminooctanoic acid lactam, cyclooctanone isooxime)** [935-30-8] **M 141.2, m 72°, 73°, 74-76°, 75°, 76-77°, b 119-122°/0.7mm, 150-151°/7-8mm, 164°/14mm, d<sub>4</sub><sup>73</sup> 1.009, n<sub>D</sub><sup>73</sup> 1.489, pK<sup>25</sup> 0.55 (AcOH).** Dissolve it in CHCl<sub>3</sub>, decolorise it with charcoal, evaporate to dryness and recrystallise it from CHCl<sub>3</sub>/hexane. Sublime it at high vacuum. The *oxime* has **m** 117° (from \*C<sub>6</sub>H<sub>6</sub> or petroleum ether). [Guggisberg *Helv Chim Acta* 61 1050 1978, Olah et al. *Synthesis* 538 1979, Behringer & Meier *Justus Liebigs Ann Chem* 607 67 1957, *Beilstein* 21 III/V 3260.]

**1*S*,2'*S*-Captopril (S-1-[3-mercapto-2-methyl-1-oxopropyl]-L-proline)** [62571-86-2] **M 217.3, m 103-104°(polymorphic unstable form m 86°, melts at 87-88° solidifies and then melts again at 104-105°), [α]<sub>D</sub><sup>22</sup> -131° (c 1.7, EtOH), pK<sub>1</sub> 3.7, pK<sub>2</sub> 9.8.** Purify it by recrystallisation from EtOAc/hexane. It is also purified by dissolving in EtOAc and chromatographed on a column of Wakogel C200 using a linear gradient of MeOH in EtOAc (0-100°) and fractions which give a positive nitroprusside test (for SH), are combined, evaporated and recrystallised from EtOAc/hexane (1:1), to give white crystals with [α]<sub>D</sub><sup>20</sup> -128.2° (c 2.0, EtOH). [Nam *J Pharm Sci* 73 1843 1984]. *Alternatively*, dissolve it in H<sub>2</sub>O, apply to a column of AG-50Wx2 (BioRad) and elute with H<sub>2</sub>O. The free acid is converted to the *dicyclohexylamine salt* in MeCN by addition of the amine until the pH is 8-9. The salt is converted to the free acid by shaking with EtOAc and 10% aqueous KHSO<sub>4</sub> or passage through an AG50Wx2 column. The EtOAc solution is dried (MgSO<sub>4</sub>), evaporated to dryness and the residue is recrystallised as above from EtOAc/hexane [Cushman et al. *Biochemistry* 16 5484 1977, NMR and IR: Horii & Watanabe *Yakugaku Zasshi (J Pharm Soc Japan)* 81 1786 1961]. It is an antihypertensive because it is a potent competitive inhibitor of the angiotensive convertive enzyme (ACE-inhibitor) with a K<sub>i</sub> value of 0.0017μM [Shimazaki et al. *Chem Pharm Bull Jpn* 30 3139 1982].

**Carbamazepine (5*H*-dibenzo[*b,f*]azepine-5-carboxamide)** [298-46-4] **M 236.3, m 190-193°, 191-192°.** Recrystallise from EtOH/\*C<sub>6</sub>H<sub>6</sub>, and dry *in vacuo*. It is soluble in Me<sub>2</sub>CO EtOH and propylene glycol, but not in H<sub>2</sub>O. Anticonvulant and Na channel inhibitor. [Stenger & Roulet *Med Exp* 11 191 1964.]

**Carbazole** [86-74-8] **M 167.2, m 240-243°, pK<sup>25</sup> <0.** Dissolve carbazole (60g) in conc H<sub>2</sub>SO<sub>4</sub> (300ml), extract with three 200ml portions of \*benzene, then stir this into 1600ml of an ice-water mixture. The precipitate is filtered off, washed with a little water, dried, recrystallised from \*benzene and then from pyridine/\*benzene [Feldman et al. *J Am Chem Soc* 73 4341 1951]. It has also been recrystallised from EtOH or toluene, sublimed in vacuum, zone-refined, and purified by TLC. [UV: Armarego in *Physical Methods in Heterocyclic Chemistry* (Ed Katritzky, Academic Press) Vol III 158 1971, *Beilstein* 20/8 V 9.]

**9-Carbazoleacetic acid** [524-80-1] **M 225.2, m 215°, pK<sub>Est</sub> ~3.5.** Crystallise the acid from ethyl acetate. [Fan et al. *Anal Chim Acta* 367 1 1998, *Beilstein* 20 III/IV 3841.]

**Carbazole-9-carbonyl chloride** [73500-82-0] **M 300.0, m 100-103°, 103.5-104.5°.** Recrystallise the acid chloride from \*C<sub>6</sub>H<sub>6</sub>. If it is not very pure (presence of OH or NH bands in the IR), dissolve it in pyridine, shake it with phosgene in toluene, evaporate and recrystallise the residue. Carry out this experiment in a good

fume cupboard as  $\text{COCl}_2$  is very **TOXIC**, and store the product in the dark. It is moisture sensitive. The *amide* has **m** 246.5-247°, and the *dimethylaminoethylamide hydrochloride* has **m** 197-198°. [Weston et al. *J Am Chem Soc* **75** 4006 1953, *Beilstein* **20** III/IV 3841.]

**1-Carbethoxy-4-methylpiperazine hydrochloride** [532-78-5] **M** 204.7, **m** 168.5-169°, **pK**<sup>25</sup> 7.31. Crystallise the hydrochloride from absolute EtOH. Its solubility is 280g/100g of  $\text{H}_2\text{O}$  at 25°. The *free base ester* has **b** 97-98°/8mm. [*Beilstein* **23** III/IV 223.]

**$\gamma$ -Carboline (9H-pyrido[3,4-b]indole)** [244-69-9] **M** 168.2, **m** 225°, 230-233°, **pK**<sup>25</sup> ~0. Crystallise it from water or EtOAc. [Robinson & Thornley *J Chem Soc* **12S** 2169 1924, Dalton et al. *Aust J Chem* **22** 185 1969, Staab & Wendel *Org Synth* **48** 44 1968, *Beilstein* **23** II 223, **23** III/IV 1572.]

***N,N'*-Carbonyldiimidazole** [530-62-1] **M** 162.2, **m** 115.5-116°. Crystallise it from \*benzene or tetrahydrofuran in a dry-box and store it dry. [Hearn *Methods Enzymol* **135** 102 1987, *Beilstein* **23/4** V 245.]

**1,1'-Carbonyldi(1,2,4-triazole) (CDT)** [41864-22-6] **M** 164.1, **m** 134-136°, 145-150°. Dissolve CDT in tetrahydrofuran and evaporate at 10mm until it crystallises. Wash the crystals with cold tetrahydrofuran and dry them in a vacuum desiccator over  $\text{P}_2\text{O}_5$  in which it can be stored for months. [Bergmann & van den Brink *Rec Trav Chim Pays Bas* **80** 1372 1961, Potts *J Org Chem* **27** 2631 1962, Staab *Justus Liebigs Ann Chem* **106** 75 1957.]

**(±)-Catechin** [7295-85-4] **M** 272.3, **m** 177° (anhydrous). Crystallise it from hot water and dry it at 100°. [*Beilstein* **17/8** V 448.]

**Cetylpyridinium chloride ( $\text{H}_2\text{O}$ ) (hexadecylpyridinium chloride)** [6004-24-6] **M** 358.0, **m** 80-83°. Crystallise the chloride from MeOH or EtOH/diethyl ether and dry it *in vacuo*. [Moss et al. *J Am Chem Soc* **108** 788 1986, Lennox & McClelland *J Am Chem Soc* **108** 3771 1986, *Beilstein* **20** V 233.]

**Chelerythrine (1,2-dimethoxy-12-methyl[1,3]benzodioxolo[5,6-c]phenanthridinium)** [34316-15-9] **M** 389.4, **m** 207°. Chelerythrine crystallises from  $\text{CHCl}_3$  on addition of MeOH [Manske *Can J Res* **21B** 140 1943, UV: Hruban et al. *Col Czech Chem Commun* **35** 3420 1970]. The *pseudo base* is colourless while the salts are yellow in aqueous solution and are fluorescent.

**Chelidonic acid (4-oxopyran-2,6-dicarboxylic acid)** [99-32-1] **M** 184.1, **m** 262°, **pK**<sub>2</sub><sup>25</sup> 2.36. The acid crystallises from aqueous EtOH. Dry it first at 100°/2 hours, then at 160° to constant weight to remove water of crystallisation. It decarboxylates at 220-230° in a vacuum. [Riegel & Zwilmeyer *Org Synth Coll Vol* **II** 126 1943, *Beilstein* **18** H 490, **18/8** V 646.]

**2-Chlorobenzothiazole** [615-20-3] **M** 169.6, **m** 21°, 90-91.4°/4mm, 135-136°/28mm, **d**<sub>4</sub><sup>20</sup> 1.303, **n**<sub>D</sub><sup>20</sup> 1.6398. It is purified by fractional distillation *in vacuo*. The *2-chloro-3-methylbenzothiazolinium 2,4-dinitrobenzenesulfonate* crystallises from  $\text{Ac}_2\text{O}$ , **m** 162-163°(dec). [Young & Amstutz *J Am Chem Soc* **73** 4773 1951, Brower et al. *J Org Chem* **19** 1830 1954, Hunter & Jones *J Chem Soc* 2190 1930, *Beilstein* **27** H 44, **27** II 18, **27** III/IV 1072.]

***o*-Chlorobenzotrifluoride** [88-16-4] **M** 180.6, **b** 152.3°, **d**<sub>4</sub><sup>20</sup> 1.371, **n**<sub>D</sub><sup>20</sup> 1.456. Dry it over  $\text{CaSO}_4$ , and distil it at high reflux ratio. [*Beilstein* **5** III 692.]

**2-Chlorobenzoxazole** [615-18-9] **M** 153.6, **b** 95-96°/20mm, 198-202°/atm, **d**<sub>4</sub><sup>20</sup> 1.331, **n**<sub>D</sub><sup>20</sup> 1.570. Purify it by fractional distillation, preferably in a vacuum. [Siedel *J Prakt Chem* **42** 456 1890, Katz *J Am Chem Soc* **75** 712 1953, Meyer & Sigel *J Org Chem* **42** 2769 1977, *Beilstein* **27** H 43, **27** II 17.]

**2-(4-Chlorobutyl)-1,3-dioxolane** [118336-86-0] **M** 164.6, **b** 56-58°/0.1mm, **d**<sub>4</sub><sup>20</sup> 1.106, **n**<sub>D</sub><sup>20</sup> 1.457. If the IR has no CHO band, then just distil it in a vacuum. If it is present, then dissolve it in  $\text{Et}_2\text{O}$ , wash it with  $\text{H}_2\text{O}$ , then saturated  $\text{NaHCO}_3$ , dry over  $\text{MgSO}_4$ , evaporate and distil it. [cf Kriesat & Gisvold *J Pharm Sci* **60** 1250

1971, Loftfield *J Am Chem Soc* **73** 1365 1951.]

**4-Chloro-2,6-diaminopyrimidine** (2,4-diamino-6-chloropyrimidine) [156-83-2] **M 144.6, m 198°, 199-202°, pK<sup>25</sup> 3.57**. It recrystallises from boiling H<sub>2</sub>O (charcoal) as needles; it also crystallises from Me<sub>2</sub>CO. [Büttner *Chem Ber* **36** 2232 1903, Roth *J Am Chem Soc* **72** 1914 1950, UV: Brown & Jacobsen *J Chem Soc* 3172 1962, *Beilstein* **24** H 318, **25** III/IV 2788.]

**2-Chloro-4,6-dimethylpyrimidine** [4472-44-0] **M 142.6, m 34-38°, 38°, b 223°/atm, pK<sup>20</sup> -0.68**. The chloro-pyrimidine has been distilled at atmospheric pressure, and solidifies on cooling. Purify it by recrystallisation from petroleum ether (b 40-60°). The 2-chloro-substituent is readily subject to nucleophilic substitution, and the kinetics of reaction with morpholine and piperidine are reported in detail (cf Chapman & Rees). It is a very weak base and requires 2N HCl to form the cation species. The UV has  $\lambda_{\max}$  (log $\epsilon$ ) at 255 (3.65) and 263 (infl 3.49) nm at pH 2.0 (neutral species); and 261nm (3.87) in ~2N HCl (H<sub>0</sub> -3.0, cation). [Boarland & McOmie *J Chem Soc* 1218 1952, Boarland & McOmie 3722 1952, Chapman & Rees *J Chem Soc* 1190 1954, Brown & England *J Chem Soc, C* 1922 1967, Matsukawa & Ohta *J Pharm Soc, Jpn* **69** 491 1949, *Chem Abstr* **44** 3456 1950, Andrisano & Modena *Gazz Chem Ital* **81** 405 1951, Angerstein *Chem Ber* **34** 3956 1901, *Beilstein* **23** H 95, **23** IV 918.]

**2-Chloro-3,5-dinitropyridine** [2578-45-2] **M 203.5, m 62-65°, 63-65°, 64°, pK<sub>Est</sub> <-5**. Dissolve it in CHCl<sub>3</sub>, shake it with saturated NaHCO<sub>3</sub>, dry (MgSO<sub>4</sub>), evaporate and apply to an Al<sub>2</sub>O<sub>3</sub> column, elute with petroleum ether (b 60-80°), evaporate and recrystallise it from \*C<sub>6</sub>H<sub>6</sub> or petroleum ether. [Ochiai & Kaneko *Chem Pharm Bull Jpn* **8** 28 1960, Plazek *Rec Trav Chim Pays Bas* **72** 573 1953, *Beilstein* **20/5** V 458.]

**1-(2-Chloroethyl)pyrrolidine hydrochloride** [7250-67-1] **M 170.1, m 167-170°, 173.5-174°, pK<sub>Est</sub> ~8.5 (free base)**. Purify the hydrochloride by recrystallisation from isopropanol/di-isopropyl ether (charcoal) and recrystallise it twice more. The *free base*, **b 55-56°/11mm, 60-63°/23mm and 90°/56mm**, is relatively unstable and should be converted to the hydrochloride immediately, by dissolving in isopropanol and bubbling dry HCl through the solution at 0°, filtering off the hydrochloride and recrystallising it. The *picrate* has **m 107.3-107.8°** (from EtOH) [Cason *J Org Chem* **24** 247 1959, Wright et al. *J Am Chem Soc* **70** 3098 1948]. [*Beilstein* **20** III/IV 66.]

**5-Chloroindole** [17422-32-1] **M 151.6, m 69-71°, 72-73°, b 120-130°/0.4mm, pK<sub>Est</sub> <0**. It is distilled at high vacuum and recrystallises from petroleum ether (b 40-60°) or (b 80-100°) as glistening plates. The *picrate* has **m 147° (146.5-147.5°)**(from \*C<sub>6</sub>H<sub>6</sub>). [Rydon & Tweddle *J Chem Soc* 3499 1955, Sugasawa *J Org Chem* **44** 578 1979, *Beilstein* **20/4** V 34.]

**2-Chloro-6-(methylamino)purine** [82499-02-3] **M 183.6, m >220°, > m 300°, pK<sub>Est(1)</sub> ~3.0, pK<sub>Est(2)</sub> ~10.0**. Purify by recrystallisation from glacial acetic acid or \*C<sub>6</sub>H<sub>6</sub>/EtOCH<sub>2</sub>CH<sub>2</sub>OH, and dry *in vacuo*. Its UV has  $\lambda_{\max}$ ( $\epsilon$ ) at 267(14400) (0.1N HCl), 271(15000) (aqueous pH 7.0); and 226(15300) and 272(5340) (0.1N NaOH in EtOH) [Montgomery & Holum *J Am Chem Soc* **80** 404 1958]. [Kim et al. *J Med Chem* **43** 746 2000, *Beilstein* **26** III/IV 3724.]

**2-Chloro-3-methylindole (2-chloroskatole)** [51206-73-6] **M 165.6, m 114.5-115.5°, pK<sub>Est</sub> <0**. Purify 2-chloroskatole by chromatography on silica gel in CH<sub>2</sub>Cl<sub>2</sub>/petroleum ether (1:2), followed by recrystallisation from aqueous EtOH, aqueous AcOH or petroleum ether (**m 113.5°**). The *picrate* has **m 121-122°(dec)** (orange-red crystals from petroleum ether). [Phillips & Cohen *J Am Chem Soc* **108** 2023 1986, *Beilstein* **20** H 317, **20** III/IV 3212.]

**4-(Chloromethyl)pyridine hydrochloride** [1822-51-1] **M 164.0, m 170-175°, 172-173°, pK<sub>Est</sub> ~5.6**. Purify it by recrystallisation from EtOH or EtOH/dry Et<sub>2</sub>O. It melts between 171° and 175°, and the clear melt resolidifies on further heating at 190° and turns red to black at 280° but does not melt again. The *picrate-hydrochloride* (prepared in EtOH) has **m 146-147°**. The *free base* is an oil. [Mosher & Tessieri *J Am Chem Soc* **73** 4925 1951, *Beilstein* **20** III/IV 2752.]

**2-Chloro-1-methylpyridinium iodide** [14338-32-0] M 255.5, m 203-205°, 205-206°(dec), 207°. Purify it by dissolving in EtOH and adding dry Et<sub>2</sub>O. The solid is washed with Me<sub>2</sub>CO and dried at 20°/0.35mm. Store it in the dark. Attempted recrystallisation from Me<sub>2</sub>CO/EtOH/petroleum ether (b 40-60°) causes some exchange of the Cl substituent by I. The *picrate* has m 106-107°, and the *perchlorate* has m 212-213°. [Jones et al. *J Am Chem Soc* **111** 1157 1989, UV and solvolysis: Barlin & Benbow *J Chem Soc, Perkin Trans 2* 790 1974, *Beilstein* **20/5** V 405.]

**6-Chloronicotinic acid** [5326-23-8] M 157.6, m 190-193°, 198-199°(dec), pK<sup>25</sup> 4.22 (50% aqueous EtOH). Purify it by recrystallisation from hot H<sub>2</sub>O and sublime it in a vacuum. [Pechmann & Welsch *Chem Ber* **17** 2384 1884, Herz & Murty *J Org Chem* **26** 122 1961, *Beilstein* **22/2** V 177.]

**4-Chloro-7-nitrobenzofurazane (7-chloro-4-nitrobenzoxadiazole, NBD-chloride)** [10199-89-0] M 199.6, m 96.5-97°, 97°, 99-100°. Wash the solid with H<sub>2</sub>O, and it recrystallises from aqueous EtOH (1:1) as pale yellow needles. It sublimes in a vacuum [Gosh & Whitehouse *Biochem J* **108** 155 1968, UV, NMR: Bolton et al. *J Chem Soc* 1004 1966].

**2-Chloro-3-nitropyridine** [5470-18-8] M 158.5, m 100-103°, 101-102°, 103-104° (sublimes), pK<sup>20</sup> -2.6. It forms needles from H<sub>2</sub>O. Purify it by continuous sublimation over a period of 2 weeks at 50-60°/0.1mm [Barlin *J Chem Soc* 2150 1964]. The *N-oxide* has m 100°(from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O). [Taylor & Driscoll *J Org Chem* **25** 1716 1960, Ochiai & Kaneko *Chem Pharm Bull Jpn* **8** 28 1960, *Beilstein* **20/5** V 451.]

**2-Chloro-5-nitropyridine** [4548-45-2] M 158.5, m 108°, pK<sub>Est</sub> ~-2.6. It crystallises from \*benzene or \*benzene/petroleum ether. [*Beilstein* **20/5** V 452.]

**9-Chloro-9-phenylxanthene (Pixyl chloride)** [42506-03-6] M 292.8, m 105-106°. A possible impurity is 9-hydroxy-9-phenylxanthene. If the material contains a lot of the hydroxy product, then boil 10g of it in CHCl<sub>3</sub> (50ml) with redistilled acetyl chloride (1ml) until liberation of HCl is complete. Evaporation leaves the chlorophenylxanthene as the hydrochloride which on heating with \*benzene loses HCl; and on adding petroleum ether prisms of chlorophenylxanthene separate and contain 0.5mol of \*benzene. The \*benzene-free compound is obtained on drying, and it melts to a colourless liquid. [Gomberg & Cone *Justus Liebigs Ann Chem* **370** 142 1909.] The 9-phenylxanthyl group is called "pixyl" and is a good protecting group [Chattopadhyaya & Reese *J Chem Soc, Chem Commun* 639 1978, *Beilstein* **17** III/IV 1704.]

**Chlorophylls a and b** See entries in "Miscellaneous Compounds", Chapter 7.

**6-Chloropurine** [87-42-3] M 154.6, m 179°(dec), pK<sub>1</sub><sup>20</sup> 0.45, pK<sub>2</sub><sup>20</sup> 7.88. 6-Chloropurine crystallises from water. The UV in water at pH 1 has λ<sub>max</sub> 264nm (log<sub>e</sub> 3.94). [*Beilstein* **26** III/IV 1742.]

**2-Chloropyrazine** [14508-49-7] M 114.5, b 62-63°/31mm, 153-154°/atm, d<sub>4</sub><sup>20</sup> 1.302, n<sub>D</sub><sup>24</sup> 1.535, pK<sub>Est</sub> <0. Fractionally distil it through a short column packed with glass helices. It has a penetrating, mildly pungent odour with a high vapour pressure at room temperature. [Erickson & Spoerri *J Am Chem Soc* **68** 400 1946, Hetman & O'Donnell *J Org Chem* **28** 1682 1963, *Beilstein* **23/5** V 366.]

**2-Chloropyridine** [109-09-1] M 113.6, b 49.0°/7mm, d<sub>4</sub><sup>20</sup> 1.20, n<sub>D</sub><sup>20</sup> 1.532, pK<sup>20</sup> 0.49 (0.72). Dry 2-chloropyridine with NaOH for several days, then distil it from CaO under reduced pressure. [*Beilstein* **20/5** V 402.]

**3-Chloropyridine** [626-60-8] M 113.6, b 148°, d<sub>4</sub><sup>20</sup> 1.194, n<sub>D</sub><sup>20</sup> 1.533, pK<sup>20</sup> 2.84. Distil 3-chloropyridine from KOH pellets. [*Beilstein* **20/5** V 406.]

**4-Chloropyridine** [626-61-9] M 113.6, b 85-86°/100mm, 147-148°/760mm, pK<sup>20</sup> 3.84. Pour 4-chloropyridine into distilled water, and excess of 6M NaOH is added to give pH 12. The organic phase is separated and extracted with four volumes of diethyl ether. The combined extracts are filtered through paper to

remove water, and the solvent is evaporated. The dark brown residual liquid is kept under high vacuum [Vaidya & Mathias *J Am Chem Soc* **108** 5514 1986]. It can be distilled, but readily darkens and is best kept as the hydrochloride [7379-35-3] **M** 150.1, **m** 163-165°(dec). [*Beilstein* **20/5** V 410.]

**2-Chloropyrimidine** [1722-12-9] **M** 114.5, **m** 63-65°, 66°, **b** 91°/26mm, **pK**<sup>20</sup> -1.90. It has been recrystallised from \*C<sub>6</sub>H<sub>6</sub>, petroleum ether or a mixture of both. It sublimes at 50°/18mm and can be distilled in a vacuum. [IR: Short & Thompson *J Chem Soc* 168 1952, Boarland & McOmie *J Chem Soc* 1218 1951, *Beilstein* **23/5** V 343.]

**2-Chloroquinoline** [612-62-4] **M** 163.6, **m** 34°, **b** 147-148°/15mm, **d**<sub>4</sub><sup>35</sup> 1.235, **n**<sub>D</sub><sup>25</sup> 1.629, **pK**<sub>Est</sub> ~0.3. Purify it by crystallisation of its *picrate* to constant melting point (123-124°) from \*benzene, regenerating the base and distilling it under vacuum [Cumper et al. *J Chem Soc* 1183 1962]. 2-Chloroquinoline can be crystallised from EtOH. Its *picrate* has **m** 123-124° (from EtOH). [*Beilstein* **20** H 359, **20/7** V 312.]

**4-Chloroquinoline** [611-35-8] **M** 163.6, **m** 29-32°, 31°, **b** 130°/15mm, 261°/744mm, **pK**<sup>25</sup> 3.72. Possible impurities include the 2-isomer. It is best purified by converting to the *picrate* (**m** 212-213° dec) in EtOH and recrystallising it from EtOH (where the *picrate* of the 2-chloroquinoline remains in solution) or EtOAc. The *picrate* is decomposed with 5% aqueous NaOH, extracted in CHCl<sub>3</sub>, washed with H<sub>2</sub>O, dried (MgSO<sub>4</sub>), evaporated and distilled in a vacuum. It can be steam distilled from slightly alkaline aqueous solutions, the aqueous distillate is extracted with Et<sub>2</sub>O, evaporated and distilled. The distillate solidifies on cooling. [Bobrński *Chem Ber* **71** 578 1938, *Beilstein* **20/7** V 314.]

**8-Chloroquinoline** [611-33-6] **M** 163.6, **b** 171-171.5°/24mm, **d**<sub>4</sub><sup>20</sup> 1.278, **n**<sub>D</sub><sup>20</sup> 1.644, **pK**<sup>20</sup> 3.12. Purify it by crystallisation of its ZnCl<sub>2</sub> complex (**m** 228°) from aqueous EtOH. [*Beilstein* **20** III/IV 3381, **20/7** V 315.]

**5-Chloroquinolin-8-yl trifluoromethanesulfonate (5-chloro-8-quinoline triflate)** [157437-38-2] **M** 311.7, **m** 79-83°. If it is discoloured dissolve it in CH<sub>2</sub>Cl<sub>2</sub>, wash it with N NaOH and half saturated K<sub>2</sub>CO<sub>3</sub>, dry the organic layer over solid K<sub>2</sub>CO<sub>3</sub>, filter, evaporate and dry the solid *in vacuo*, (cf 2-methyl-5-pyridine triflate). [Matthew et al. *Tetrahedron Lett* **35** 5177 1994, Tilley & Zawoiski *J Org Chem* **53** 386 1988, Ellingboe et al. *J Med Chem* **37** 542 1994.]

**8-Chlorotheophylline (8-chloro-1,3-dimethyl-2,6(1H,3H)-purinedione)** [85-18-7] **M** 214.6, **m** 311°(dec), **pK**<sub>Est(1)</sub> ~5.4, **pK**<sub>Est(2)</sub> ~9.1. It crystallises from H<sub>2</sub>O or EtOH (**m** 304°, dec). The *choline salt* crystallises from H<sub>2</sub>O with **m** 60-62° (2 H<sub>2</sub>O) and **m** 97-99° (anhydrous). [*Beilstein* **26** H 473, II 276, **26** III/IV 2442.]

**2-Chlorothiophene (2-thienyl chloride)** [96-43-5] **M** 118.6, **b** 126-128°, 128°/~760mm, **d**<sub>4</sub><sup>20</sup> 1.285, **n**<sub>D</sub><sup>20</sup> 1.551. Purify it by fractional distillation at atmospheric pressure or by gas chromatography. [Conde et al. *Synthesis* 412 1976, *Beilstein* **17/1** V 303.]

**5-Chlorouracil (5-chloro-2,4(6)-dihydropyrimidine)** [1820-81-1] **M** 146.5, **m** 314-418°(dec), 324-325°(dec), **pK**<sub>1</sub><sup>25</sup> 7.95, **pK**<sub>2</sub><sup>25</sup> >13. It recrystallises from hot H<sub>2</sub>O (4g/500ml) using charcoal. [McOmie et al. *J Chem Soc* 3478 1955, West & Barrett *J Am Chem Soc* **76** 3146 1954, *Beilstein* **24** III/IV 1231.]

**4-Chromanone (2,3-dihydro-4H-1-benzopyran-4-one)** [491-37-2] **M** 148.2, **m** 35-37°, 39°, 41°, **b** 92-93°/3mm, 130-132°/15mm, 160°/50mm. It has been recrystallised from petroleum ether, or purified by dissolving in \*C<sub>6</sub>H<sub>6</sub> washing with H<sub>2</sub>O, drying (MgSO<sub>4</sub>), evaporating and distilling in a vacuum, then recrystallising the residue. The liquid has a pleasant lemon-like odour. The *semicarbazone* has **m** 227°. [Loudon & Razdan *J Chem Soc* 4299 1954.] The *oxime* is prepared from 3g of chromanone, 3g NH<sub>2</sub>OH.HCl in EtOH (50ml), 6g K<sub>2</sub>CO<sub>3</sub> and refluxed on a water bath for 6 hours. The solution is poured into H<sub>2</sub>O, the solid is filtered off, dried and dissolved in hot \*C<sub>6</sub>H<sub>6</sub> which on addition of petroleum ether yields the *oxime* as glistening needles **m** 140°. Hydrolysis of this gives very pure chromanone. The *benzal derivative* is prepared from 3g of chromanone and 4g PhCHO in 50ml EtOH, heated to boiling, 10ml of conc HCl are added dropwise and set aside for several days. The derivative separates and is recrystallised from EtOH to give yellow needles, **m** 112° [Powell *J Am Chem Soc* **45** 2711 1923]. Reaction with Pb(OAc)<sub>4</sub> yields the *3-acetoxy derivative* **m** 74°

(from petroleum ether + trace of EtOAc) [Cavill et al. *J Chem Soc* 4573 1954, *Beilstein* 17/10V 14].

**Cinnoline** [253-66-7] **M** 130.2, **m** 38°, 40-41°, **b** 114°/0.35, **pK**<sup>20</sup> 2.37. It is distilled at high vacuum, then recrystallised from petroleum ether. Keep it under N<sub>2</sub> in sealed tubes in the dark at 0°. The *hydrochloride* [5949-24-6] **M** 166.6 crystallises from EtOH/Et<sub>2</sub>O and has **m** 156-158° (yellow *monohydrate*), and dehydrates on sublimation at 110-115°/3mm. The *picrate* has **m** 196-196.5°. [*Beilstein* 23 H 173, 23 III.IV 1217.]

**Citrazinic acid (2,6-dihydroxyisonicotinic acid)** [99-11-6] **M** 155.1, **m** >300°, **pK**<sub>1</sub> 3.0, **pK**<sub>2</sub> 4.76. The acid is normally a yellow powder with a greenish shade, but is white when ultra pure and turns blue on long standing. It is insoluble in H<sub>2</sub>O but slightly soluble in hot HCl and soluble in alkali or carbonate solutions. It is purified by precipitation from alkaline solutions with dilute HCl, and dried in a vacuum over P<sub>2</sub>O<sub>5</sub>. The *ethyl ester* has **m** 232° (evacuated tube) and a pKa of 4.81 in MeOCH<sub>2</sub>CH<sub>2</sub>OH [IR: Pitha *Col Czech Chem Comm* 28 1408 1963]. [*Beilstein* 22/7 V 24.]

**Clioquinol (5-chloro-8-hydroxy-7-iodoquinoline)** [130-26-7] **M** 305.5, **m** 181°, **pK**<sub>1</sub><sup>25</sup> 2.7, **pK**<sub>2</sub><sup>25</sup> 7.9. It crystallises from AcOH or xylene and dry it at 70° *in vacuo*. [*Beilstein* 21 III/IV 1190.]

**Clofazimine [2-(4-chloroanilino)-3-isopropylimino-5-(4-chlorophenyl)-3,4-dihydrophenazine]** [2030-63-9] **M** 473.5, **m** 210-212°, **pK**<sup>20</sup> 8.37 (8.51). Clofazimine recrystallises from acetone as dark red crystals. Its solubility in CHCl<sub>3</sub> and EtOH is 7% and 0.1%, respectively, at room temperature. It is insoluble in H<sub>2</sub>O. It is antibacterial. [Barry et al. *J Chem Soc* 859 1958, *Beilstein* 25 III/IV 3033.]

**Colforsin**, see **Forskolin**.

**Conessine** [546-06-5] **M** 356.6, **m** 125°, 127-128.5°, [α]<sub>D</sub><sup>20</sup> -1.9° (in CHCl<sub>3</sub>) and +26° (c 3, EtOH), **pK**<sub>Est(1)</sub> ~10.4, **pK**<sub>Est(2)</sub> ~10.7. It crystallises from acetone, sublimates at 95°/0.01mm and boils at 0.1mm with bath temperature at 220°. The *dihydrochloride* has **m** >340° (browns at 235° and decomposes at 338-240°) and has [α]<sub>D</sub><sup>20</sup> +9.3° (c 2, H<sub>2</sub>O). [Marshall & Johnson *J Am Chem Soc* 84 1458 1962, *Beilstein* 22 III/IV 4382.]

**Coproporphyrin I** [531-14-6] **M** 654.7, **λ**<sub>max</sub> 591, 548, 401nm in 10% HCl. It crystallises from pyridine/glacial acetic acid. The *dihydrochloride* [69477-27-6] has **M** 727.7 and **λ**<sub>max</sub> at 395nm in water. [*Beilstein* 26 III/IV 3094.]

**Coumalic acid (2-pyrone-5-carboxylic acid)** [500-05-0] **M** 140.1, **m** 205-210°(dec), **pK**<sub>Est</sub> ~0. The acid crystallises from MeOH. The *methyl ester* has **m** 73-74° (from petroleum ether) and **b** 178-180°/60 mm. [*Beilstein* 18/8 V 120.]

**Coumarin** [91-64-5] **M** 146.2, **m** 68-69°, **b** 298°, **pK**<sup>25</sup> -4.97 (aqueous H<sub>2</sub>SO<sub>4</sub>). Coumarin crystallises from ethanol or water and sublimates *in vacuo* at 43° [Srinivasan & deLevie *J Phys Chem* 91 2904 1987]. [*Beilstein* 17/10 V 143.]

**Coumarin-3-carboxylic acid** [531-81-7] **M** 190.2, **m** 188°(dec), **pK**<sub>Est</sub> ~1.5. The acid crystallises from water. [*Beilstein* 18/8 V 323.]

**γ-Crotonolactone [2(5H)-furanone]** [497-23-4] **M** 84.1, **m** 3-4°, 76-77°/3.5mm, 90.5-91°/11.5mm, 92-93°/14mm, 107-109°/24mm, 212-214°/760mm, **d**<sub>4</sub><sup>20</sup> 1.197, **n**<sub>D</sub><sup>20</sup> 1.470. Fractionally distil the lactone under reduced pressure. Its IR(CCl<sub>4</sub>) has 1784 and 1742 cm<sup>-1</sup>, UV no max above 205nm (ε 1160 cm<sup>-1</sup> M<sup>-1</sup>) and <sup>1</sup>HNMR (CCl<sub>3</sub>) has τ at 2.15 (pair of triplets 1H), 3.85 (pair of triplets 1H) and 5.03 (triplet 2H). [Price & Judge *Org Synth Coll Vol V* 255 1973, Smith & Jones *Can J Chem* 37 2007, 2092 1959, *Beilstein* 17/9 V 112.]

**15-Crown-5** [33100-27-5] **M** 220.3, **b** 93-96°/0.1mm, **d**<sub>4</sub><sup>20</sup> 1.113, **n**<sub>D</sub><sup>20</sup> 1.465. Dry it over 3A molecular sieves and distil it in a high vacuum. [*Beilstein* 19/12 V 252.]

**18-Crown-6** [17455-13-9] **M** 264.3, **m** 37-39°. Recrystallise it from acetonitrile and dry it in a vacuum.

Purify it also by precipitating the 18-crown-6/nitromethane 1:2 complex with Et<sub>2</sub>O/nitromethane (10:1 mixture). The complex is decomposed in vacuum whereby 18-crown-6 distills off under the reduced pressure. [*Beilstein* 19/12 V 601.]

**Cryptopine** [482-74-6] M 369.4, m 220-221°, 220-223°, pK<sup>25</sup> 8.09. It crystallises from \*benzene, hot EtOH (0.25% cold, 1.2% at boiling), petroleum ether or methyl ethyl ketone. The *perchlorate* crystallises from aqueous MeOH with m 226-228°(dec). [Thomas et al. *Can J Chem* 33 570 1955, Howarth & Perkin *J Chem Soc* 1769 1926, *Beilstein* 27 III/IV 6652.]

**Cupreine (6'-hydroxycinchonidine)** [524-63-0] M 310.4, m 202°(anhydrous), [α]<sub>D</sub><sup>17</sup> -176° (c 0.5, MeOH), pK<sup>15</sup> 7.63. Cupreine crystallises from EtOH (*anhydrous* crystals) and wet Et<sub>2</sub>O (as *dihydrate* crystals). It has K<sub>b</sub> 2.7×10<sup>-7</sup> [Kolthoff *Biochem Z* 162 323]. The *sulfate* forms needles, m 257°(dec), from MeOH, amyl alcohol or H<sub>2</sub>O, with [α]<sub>D</sub><sup>20</sup> -197.9° (c 1.2, H<sub>2</sub>O). [*Beilstein* 22 I 165, 22 II 416.]

**5-Cyanoindole** [15861-24-2] M 142.2, m 106-108°, 107-108°, pK<sup>25</sup> <0. Dissolve the nitrile in 95% EtOH, boil it in the presence of charcoal, filter, evaporate to a small volume and add enough H<sub>2</sub>O to cause crystallisation and cool. Recrystallise it directly from aqueous EtOH and dry it in a vacuum. Its UV has λ<sub>max</sub> at 276 nm (log ε 3.6) in MeOH. [Lindwall & Mantell *J Org Chem* 18 345 1953, 20 1458 1955, Thesing at al. *Chem Ber* 95 2205 1962, NMR: Lallemand & Bernath *Bull Soc Chim Fr* 4091 1970, *Beilstein* 22/3 V 45.]

**3-Cyanopyridine** [100-54-9] M 104.1, m 50°, pK<sup>25</sup> 1.38. It is recrystallised to constant melting point from *o*-xylene/hexane. [*Beilstein* 22/2 V 115.]

**4-Cyanopyridine** [100-48-1] M 104.1, m 76-79°, pK<sup>25</sup> 1.86. It crystallises from dichloromethane/diethyl ether mixture. [*Beilstein* 22/2 V 214.]

**Cyanuric acid (2,4,6-trichloro-1,3,5-triazine)** [108-80-5] M 120.1, m >300°, pK<sup>25</sup> 6.78. It crystallises from water. Dry it at room temperature in a desiccator in a vacuum. [*Beilstein* 26 III/IV 632.]

**Cyanuric chloride (TCT, 2,4,6-trichloro-1,3,5-triazine)** [108-77-0] M 184.4, m 146-149°, 154°, b 190°. TCT crystallises from CCl<sub>4</sub> or petroleum ether (b 90-100°) and is dried under vacuum. It has also been recrystallised twice from anhydrous \*benzene immediately before use [Abuchowski et al. *J Biol Chem* 252 3582 1977]. [*Beilstein* 26 III/IV 66.]

**Cycloheximide (Actidione)** [66-81-9] M 281.4, m 119.5-121°, [α]<sub>D</sub><sup>20</sup> +9.5° (c 2, H<sub>2</sub>O). It crystallises from H<sub>2</sub>O/MeOH (4:1), amyl acetate, isopropyl acetate/isopropyl ether or H<sub>2</sub>O. [*Beilstein* 21/13 V 434.]

**1-Cyclohexyl-5-methyl-1H-tetrazole** [7707-57-5] M 166.2, m 124-124.5°. Crystallise it from absolute EtOH or H<sub>2</sub>O (heavy needles), then sublime it at 115°/3mm. [Harvill et al. *J Org Chem* 15 662, 668 1950, Billuber Inc USP 2507337 1946, *Beilstein* 26 III/IV 1661.]

**Cyclotrimethylenetrinitramine (RDX, Cyclonite, 1,3,5-trinitrohexahydro-1,3,5-triazine)** [121-82-4] M 222.2, m 203.8°(dec), 205-206°(dec). RDX crystallises from acetone. [Bachmann & Sheehan *J Am Chem Soc* 71 1842 1949, *Beilstein* 26 II 5, 26 III/IV 22.] **EXPLOSIVE.**

**Cytisine** (7*R*,9*S*-7,9,10,11,12,13-hexahydro-7,9-methano-12*H*-pyrido[1,2-*a*][1,5]diazocin-8-one, Laburnine, Ulexine) [485-35-8] M 190.3, m 152-153°, 155°, b 218°/2mm, [α]<sub>D</sub><sup>17</sup> -120° (H<sub>2</sub>O), [α]<sub>D</sub><sup>25</sup> -115° (c 1, H<sub>2</sub>O), pK<sup>15</sup> 1.20, pK<sup>15</sup> 8.12 [also stated are pK<sub>1</sub> 6.11, pK<sub>2</sub> 13.08]. Crystallise cytisine from acetone and sublime it in a vacuum. Its solubilities are: 77% (H<sub>2</sub>O), 7.7% (Me<sub>2</sub>CO), 28.6% (EtOH), 3.3% (\*C<sub>6</sub>H<sub>6</sub>), 50% (CHCl<sub>3</sub>) but it is insoluble in petroleum ether. The *tartrate* has m 206-207° [α]<sub>D</sub><sup>24</sup> +45.9°, the *N-tosylate* has m 206-207°, and the *N-acetate* has m 208°. [Bohlmann et al. *Angew Chem* 67 708 1955, van Tamelen & Baran *J Am Chem Soc* 77 4944 1955, Isolation: Ing *J Chem Soc* 2200 1931, Govindachari et al. *J Chem Soc* 3839 1957, Abs config: Okuda et al. *Chem Ind (London)* 1751 1961, *Beilstein* 24 H 134, 24 I 244, 24 II 70, 24 III/IV 321.] **TOXIC.**

**Cytosine (4-amino-2-hydroxypyrimidine)** See entry in “Miscellaneous Compounds”, Chapter 7.

***cis*-Decahydroisoquinoline** [2744-08-3] **M 139.2, b 97-98°/15mm, 208-209°/730mm, pK<sup>20</sup> 11.32.** The free base is treated with saturated aqueous picric acid, allowed to stand for 12 hours, filtered, washed with MeOH to remove the more soluble *trans* isomer and recrystallised from MeOH to give pure *cis-picric acid* **m** 149-150°. The picrate (~5g) is shaken with 5M aqueous NaOH (50ml) and Et<sub>2</sub>O (150ml) while H<sub>2</sub>O is added to the aqueous phase to dissolve insoluble Na picrate. The Et<sub>2</sub>O extract is dried over solid NaOH and then shaken with Al<sub>2</sub>O<sub>3</sub> (Merck for chromatography) until the yellow color of traces of picric acid disappears (this color cannot be removed by repeated shaking with 5-10 M aqueous NaOH). The extract is concentrated to 50ml and dry HCl is bubbled through until separation of the white crystals of the *cis-HCl* is complete. These are washed with Et<sub>2</sub>O, dried at 100° and recrystallised from EtOH/EtOAc to yield pure *cis-hydrochloride* **m** 182-183° (dried in a vacuum desiccator over KOH) with IR (KBr)  $\nu_{\max}$  2920, 2820, 1582, 1470, 1445, 1410, 1395, 1313, 1135, 1080, 990, 870 cm<sup>-1</sup>. The pure *free base* is prepared by dissolving the *hydrochloride* in 10 M aqueous NaOH, extracted with Et<sub>2</sub>O, dried over solid KOH, filtered and distilled in a vacuum. It has IR (film)  $\nu_{\max}$  2920, 2820, 2720, 2560, 1584, 1470, 1445, 1415, 1395, 1315, 1300, 1135, 1080, 1020, 990, 873 cm<sup>-1</sup>. The <sup>1</sup>HNMR in CDCl<sub>3</sub> is characteristically different from that of the *trans*-isomer. [Armarego *J Chem Soc (C)* 377 1967, Gray & Heitmeier *J Am Chem Soc* **80** 6274 1958, Witkop *J Am Chem Soc* **70** 2617 1948, Skita *Chem Ber* **57** 1982 1924, Helfer *Helv Chim Acta* **6** 799 1923, *Beilstein* **20** II 73, **20** III/IV 2026.]

***trans*-Decahydroisoquinoline** [2744-09-4] **M 139.2, b 106°/15mm, pK<sup>20</sup> 11.32.** This is purified as the *cis*-isomer above. The *trans-picric acid* has **m** 175-176°, and the *trans-hydrochloride* has **m** 221-222° and has IR (KBr) has  $\nu_{\max}$  at 2930, 3800, 1589, 1450, 1400, 1070, 952, 837 cm<sup>-1</sup>. The pure *free base* is prepared as above and had IR (film) with  $\nu_{\max}$  at 2920, 2820, 2720, 2560, 1584, 1470, 1445, 1415, 1395, 1315, 1300, 1135, 1080, 1020, 990, 873 cm<sup>-1</sup>. The <sup>1</sup>HNMR in CDCl<sub>3</sub> is characteristically different from that of the *cis*-isomer. (references as above and Helfer *Helv Chim Acta* **9** 818 1926). [*Beilstein* **20** II 73, **20** III/IV 2026.]

***cis*-Decahydroquinoline** [10343-99-4] **M 139.2, b 207-208°/708mm, pK<sup>20</sup> 11.29.** It is available as a *cis-trans*-mixture (**b** 70-73°/10mm, Aldrich, ~18% *cis*-isomer [2051-28-7]), but the isomers can be separated by fractionating in a spinning band column (1~1.5 metre, type E) at atmospheric pressure and collecting 2ml fractions with a distillation rate of 1 drop in 8-10 seconds. The lower boiling fraction solidifies and contains the *trans*-isomer (see below, **m** 48°). The higher boiling fraction **b** 207-208°/708mm remains liquid and is mostly pure *cis*-isomer. This is reacted with PhCOCl and M aqueous NaOH to yield the *N-benzoyl derivative* **m** 96° after recrystallisation from petroleum ether (**b** 80-100°). It is hydrolysed with 20% aqueous HCl by refluxing overnight. PhCO<sub>2</sub>H is filtered off, the filtrate is basified with 5M aqueous NaOH and extracted with Et<sub>2</sub>O. The dried extract (Na<sub>2</sub>SO<sub>4</sub>) is saturated with dry HCl gas, and the *cis-decahydroquinoline hydrochloride* which separates has **m** 222-224° after washing with Et<sub>2</sub>O and drying at 100°; and has IR (KBr) with  $\nu_{\max}$  at 2900, 2780, 2560, 1580, 1445, 1432, 1403, 1165, 1080, 1036, 990, 867 cm<sup>-1</sup>. The *free base* is obtained by dissolving the *hydrochloride* salt in 5M aqueous NaOH, extracting with Et<sub>2</sub>O and drying the extract (Na<sub>2</sub>SO<sub>4</sub>), evaporating and distilling the residue; it has IR (film) with  $\nu_{\max}$  at 2900, 2840, 2770, 1445, 1357, 1330, 1305, 1140, 1125, 1109, 1068, 844 cm<sup>-1</sup>. The <sup>1</sup>H NMR in CDCl<sub>3</sub> is characteristically different from that of the *trans*-isomer. [Armarego *J Chem Soc (C)* 377 1967, Hüchel & Stepf *Justus Liebigs Ann Chem* **453** 163 1927, Bailey & McElvain *J Am Chem Soc* **52** 4013 1930, *Beilstein* **20** H 157, **20** I 35, **20** II 72-73, **20** III/IV 2017.]

***trans*-Decahydroquinoline** [767-92-0] **M 139.2, m 48°, b 205-206°/708mm, pK<sup>20</sup> 11.29.** The lower boiling fraction from the preceding spinning band column fractionation of the commercial *cis-trans*-mixture (~ 20:60; see the *cis*-isomer above) solidifies readily (**m** 48°), and the receiver has to be kept hot with warm water. It is further purified by conversion to the *hydrochloride* **m** 285-286° after recrystallisation from EtOH/AcOEt. This has IR (KBr) with  $\nu_{\max}$  at 2920, 2760, 2578, 2520, 1580, 1455, 1070, 1050, 975, 950, 833 cm<sup>-1</sup>. The *free base* is prepared as for the *cis*-isomer above and distilled; and has IR (film, at ca 50°) with  $\nu_{\max}$  at 2905, 2840, 2780, 1447, 1335, 1305, 1240, 1177, 1125, 987, 900, 835 cm<sup>-1</sup>. The <sup>1</sup>HNMR in CDCl<sub>3</sub> is characteristically different from that of the *cis*-isomer. [Armarego *J Chem Soc (C)* 377 1967, Hüchel & Stepf *Justus Liebigs Ann Chem* **453** 163 1927, Bailey & McElvain *J Am Chem Soc* **52** 4013 1930, Prelog & Szpiflogel *Helv Chim Acta* **28** 1684 1945, *Beilstein* **20** H 157, **20** I 35, **20** II 72-73, **20** III/IV 2017.]

**Delphinine** [561-07-9] M 559.7, m 197-199°,  $[\alpha]_D^{20} +26^\circ$  (c 1, EtOH). It crystallises from EtOH with 198-200° uncorrected (187.5-188.5°) and from Et<sub>2</sub>O or Me<sub>2</sub>CO. Its solubility at ~25° in EtOH, CHCl<sub>3</sub> and Et<sub>2</sub>O is 4, 5 and 10% respectively. In c 1, EtOH, it has  $[\alpha]_D^{20} +22^\circ$  changing to +19° in 4 hours. [Markwood *J Am Chem Soc* **16** 931 1927.] The *hydrochloride* has m 214°(dec) from MeOH or MeOH/Et<sub>2</sub>O. [Weissner et al. *Can J Chem* **50** 1925 1972, Jacobs & Craig *J Biol Chem* **127** 363 1939, *Beilstein* **21** III/IV 2867.]

**3,6-Diaminoacridine hydrochloride** [952-23-8] M 245.7, m 270°(dec),  $\epsilon_{456} 4.3 \times 10^4$ ,  $pK_1^{20} 1.5$ ,  $pK_2^{20} 9.60$  (9.65 free base). It is first purified by precipitation of the free base by adding aqueous NH<sub>3</sub> solution to an aqueous solution of the hydrochloride or hydrogen sulfate (see below), drying the precipitate and subliming at 0.01mm Hg [Müller & Crothers *Eur J Biochem*, **54** 267 1975]. [*Beilstein* **22** H 487.]

**3,6-Diaminoacridine sulfate (proflavin sulfate)** [1811-28-5] M 516.6, m >300°(dec),  $\lambda_{max} 456nm$ . An aqueous solution, after treatment with charcoal, is concentrated, chilled overnight, filtered and the precipitate is rinsed with a little diethyl ether. The precipitate is dried in air, then overnight in a vacuum oven at 70°. [*Beilstein* **22** I 650, **22/11** V 323.]

**4,5-Diamino-2,6-dihydroxypyrimidine (diamino uracil) sulfate** [32014-70-3] M 382.3, m >300°,  $pK_1^{20} 1.7$ ,  $pK_2^{20} 3.20$ ,  $pK_3^{25} 4.56$ . The salt is quite insoluble in H<sub>2</sub>O but can be converted to the free base which is recrystallised from H<sub>2</sub>O and converted to the sulfate by addition of the required amount of H<sub>2</sub>SO<sub>4</sub>. The *hydrochloride* has m 300-305°(dec) and can be used to prepare the sulfate by addition of H<sub>2</sub>SO<sub>4</sub>; it is more soluble than the sulfate. The *perchlorate* has m 252-254°. The free base has  $\lambda_{max}$  at 260nm (log  $\epsilon$  4.24) in 0.1M HCl. [Bogert & Davidson *J Am Chem Soc* **86** 1668 1933, Bredereck et al. *Chem Ber* **86** 850 1953, Sherman & Taylor *Org Synth Coll Vol IV* 247 1963, Barlin & Pfeleiderer *J Chem Soc (B)* 1425 1971, *Beilstein* **25** II 382.]

**5,6-Diamino-1,3-dimethyluracil hydrate (5,6-diamino-1,3-dimethyl-2-pyrimidine-2,4-dione hydrate)** [5440-00-6] M 188.2, m 205-208°(dec), 209°(dec), 210°(dec),  $pK_1 1.7$ ,  $pK_2 4.6$ . It recrystallises from EtOH. The *hydrochloride* has m 310° (from MeOH), and the *perchlorate* has m 246-248°. [UV: Bredereck et al. *Chem Ber* **92** 583 1959, Taylor et al. *J Am Chem Soc* **77** 2243 1955, *Beilstein* **25** III/IV 4133.]

**6,9-Diamino-2-ethoxyacridine (Ethacridine)** [442-16-0] M 257.3, m 226°,  $pK^{20} 11.6$ . It crystallises from 50% EtOH or EtOH as orange-yellow crystals. It also crystallises as a *monohydrate* m 116-118°. It has a  $pK^{20}$  of 11.04 in 50% aqueous EtOH. The *methiodide* is soluble in H<sub>2</sub>O and has m 332-334° (dec) (from aqueous Me<sub>2</sub>CO). [Albert & Gledhill *J Soc Chem Ind* **61** 159 1942, Foye et al. *J Pharm Sci* **57** 1795 1968, Albert & Goldacre *J Chem Soc* 708 1946, *Beilstein* **22** II 458, **22** III/IV 6679, **22/12** V 243.]

**6,9-Diamino-2-ethoxyacridine dl-lactate monohydrate (Rivanol, Acrinol)** [6402-23-9] M 361.4, m 235° (dark at ~200°),  $pK^{20} 11.6$ . It forms yellow crystals from 90% EtOH/Et<sub>2</sub>O. Its solubility in H<sub>2</sub>O is ~15% at 25° and ~9% at 100°, and its solutions have a yellow fluorescence which is stable on boiling. It is an antiseptic. See ethacridine above, *Beilstein* **22** II 458, **22** III/IV 6680, **22/12** V 243.]

**2,4-Diamino-6-hydroxypyrimidine** [56-06-4] M 126.1, m 260-270°(dec),  $pK_1^{25} 1.34$ ,  $pK_2^{25} 3.27$ ,  $pK_3^{25} 10.83$ . It recrystallises from H<sub>2</sub>O. [*Beilstein* **25** III/IV 3642.]

**4,5-Diamino-6-hydroxypyrimidine hemisulfate** [102783-18-6] M 350.3, m 268°, 270°,  $pK_1^{25} 1.34$ ,  $pK_2^{25} 3.57$ ,  $pK_3^{25} 9.86$ . It recrystallises from H<sub>2</sub>O. The *free base* also crystallises from H<sub>2</sub>O (m 239°). [Mason *J Chem Soc* 2071 1954, Elion et al. *J Am Chem Soc* **74** 411 1952, *Beilstein* **25** III/IV 3645.]

**2,4-Diamino-5-phenylthiazole (Amiphenazole)** [490-55-1] M 191.3, m 163-164°(dec). The thiazole crystallises from aqueous EtOH or water. Store it in the dark under N<sub>2</sub>. The *hydrochloride* has m 273-274°(dec) (from MeOH/EtOAc), and the *picrate* has m 189-191°(dec) (from H<sub>2</sub>O). [Davies et al. *J Chem Soc* 3491 1950, Dodson & Turner *J Am Chem Soc* **73** 4517 1951, *Beilstein* **27** III/IV 5139.]

**2,3-Diaminopyridine** [452-58-4] M 109.1, m 116°,  $pK_1^{25} -0.50$ ,  $pK_2^{25} 6.92$ . It crystallises from \*benzene

and sublimes *in vacuo*. [Beilstein 22/11 V 241.]

**2,6-Diaminopyridine** [141-86-6] M 109.1, m 121.5° pK<sub>Est(1)</sub> <-6.0, pK<sub>Est(2)</sub> ~7.3. It crystallises from \*benzene and sublimes *in vacuo*. [Beilstein 22 III/IV 255.]

**3,4-Diaminopyridine** [54-96-6] M 109.1, m 218-219°, pK<sub>1</sub><sup>20</sup> 0.49, pK<sub>2</sub><sup>20</sup> 9.14. It crystallises from \*benzene and is stored under N<sub>2</sub> because it is *deliquescent* and absorbs CO<sub>2</sub>. [Beilstein 22/11 V 266.]

**3,5-Diamino-1,2,4-triazole (Guanazole)** [1455-77-2] M 99.1, m 206°, pK<sub>1</sub><sup>20</sup> 4.43, pK<sub>2</sub><sup>20</sup> 12.12. The triazole crystallises from water or EtOH. [Beilstein 26 III/IV 1161.]

**1,3-Diazaazulene (cycloheptimidazole)** [275-94-5] M 130.1, m 110-112°, 120°. It is recrystallised repeatedly from de-aerated cyclohexane in the dark or from petroleum ether/\*C<sub>6</sub>H<sub>6</sub> and forms yellow needles. It is soluble in H<sub>2</sub>O, EtOH and \*C<sub>6</sub>H<sub>6</sub>, and forms a *monohydrate* which loses H<sub>2</sub>O at 60°. The *picrate* has m 207°(dec). [Nozoe et al. *J Am Chem Soc* 76 3352 1954, Nukai et al. *Bull Chem Soc Jpn* 40 1967 1967, Beilstein 23 III/IV 1216.]

**1,5-Diazabicyclo[4.3.0]non-5-ene (DBN, 2,3,4,6,7,8-hexahydropyrrolo[1,2-*a*]-pyrimidine)** [3001-72-7] M 124.2, b 96-98°/11mm, 100-102°/12mm, 118-121°/32mm, d<sub>4</sub><sup>20</sup> 1.040, n<sub>D</sub><sup>20</sup> 1.520, pK<sup>25</sup> >13.0. Distil DBN from BaO. It forms a *hydroiodide* on addition of 47% HI; dry it and dissolve it in MeCN, evaporate and repeat; recrystallise from EtOH, dry at 25°/1mm for 5 hours, then at 80°/0.03mm for 12 hours and store and dispense it in a dry box, m 154-156° [Jaeger et al. *J Am Chem Soc* 101 717 1979]. The *methiodide* is recrystallised from CHCl<sub>3</sub>/Et<sub>2</sub>O, m 248-250°, and *hydrogen fumarate* has m 159-160° and is crystallised from *iso*-PrOH [Rokach et al. *J Med Chem* 22 237 1979, Oediger et al. *Chem Ber* 99 2012 1966, Reppe et al. *Justus Liebigs Ann Chem* 596 210 1955]. [Beilstein 23/5 V 239.]

**1,4-Diazabicyclo[2.2.2]octane (DABCO, triethylenediamine, TED)** [280-57-9] M 112.2, m 156-157° (sealed tube), pK<sub>1</sub><sup>25</sup> 2.97, pK<sub>2</sub><sup>25</sup> 8.82. DABCO crystallises from 95% EtOH, petroleum ether or MeOH/diethyl ether (1:1). Dry it under vacuum over CaCl<sub>2</sub> and BaO. It can be sublimed *in vacuo*, and readily at room temperature. It has also been purified by removal of water during azeotropic distillation of a \*benzene solution. It is then recrystallised twice from anhydrous diethyl ether under argon, and stored under argon [Blackstock et al. *J Org Chem* 52 1451 1987]. [Beilstein 23/3 V 487.]

**1,8-Diazabicyclo[5.4.0]undec-7-ene (DBU, 2,3,4,6,7,8,9,10-octahydropyrimidino[1,2-*a*]-azepine)** [6674-22-2] M 152.2, b 115°/11mm, d<sub>4</sub><sup>20</sup> 1.023, n<sub>D</sub><sup>20</sup> 1.522, pK<sub>Est</sub> ~ >13. Fractionally distil DBU under vacuum. Also purify it by chromatography on Kieselgel and eluting with CHCl<sub>3</sub>/EtOH/25% aqueous NH<sub>3</sub> (15:5:2) and checking by IR and MS. [Oediger et al. *Chem Ber* 99 2012 1962, *Angew Chem, Int Ed Engl* 6 76 1967, Guggisberg et al. *Helv Chim Acta* 61 1057 1978, Beilstein 23/5 V 271.]

**1,8-Diazabiphenylene** [259-84-7] M 154.2, m 156-158.5°, pK<sub>Est</sub> ~4.4. Recrystallise it from cyclohexane, then sublime it in a vacuum. [Barton & Walker *Tetrahedron Lett* 569 1975, Deroski & Marhgraf *Can J Chem* 62 2235 1984.]

**2,7-Diazabiphenylene** [31857-42-8] M 154.2, m 192-192.5°, pK<sub>Est</sub> ~4.5. It forms yellow crystals from cyclohexane, and sublimes in a vacuum. [MacBride *J Chem Soc, Chem Commun* 359 1974, Kramer & Berry *J Am Chem Soc* 94 8336 1972.]

**1,8-Diazafluorenone (cyclopenta[1.2-*b*:4,3-*b'*]dipyridin-9-one)** [54078-29-4] M 182.2, m 205°, 229-231°, pK<sub>Est</sub> ~2.6. Recrystallise it from Me<sub>2</sub>CO. The *oxime* has m 119-200°. [Druey & Schmid *Helv Chim Acta* 33 1080 1950, Beilstein 24 III/IV 622.]

**Dibenzo-18-crown-6** [14187-32-7] M 360.4, m 163-164°. Crystallise it from \*benzene, *n*-heptane or toluene and dry it under vacuum at room temperature for several days. [Szezygiel *J Phys Chem* 91 1252 1987, Vögtle ed. *Top Corr Chem (Host Guest Complex Chemistry)* 98 1981.]

**Dibenzo-18-crown-8** [14174-09-5] **M 448.5, m 103-106°**. Recrystallise it from EtOH, and dry in a vacuum at 60° over P<sub>2</sub>O<sub>5</sub> for 16 hours. [Delville et al. *J Am Chem Soc* **109** 7293 1987, Vögtle ed. *Top Corr Chem* (Host Guest Complex Chemistry) **98** 1981.]

**Dibenzofuran** [132-64-9] **M 168.2, m 82.4°**. Dissolve dibenzofuran in diethyl ether, then shake it with two portions of aqueous NaOH (2M), wash it with water, separate and dry (MgSO<sub>4</sub>) it. After evaporating the ether, dibenzofuran is crystallised from aqueous 80% EtOH and then dried under vacuum. [Cass et al. *J Chem Soc* 1406 1958.] High purity material is obtained by zone refining. [Beilstein **17** V 234.]

**Dibenzothiophene** [132-65-0] **M 184.3, m 99°**. Purify dibenzothiophene by chromatography on alumina with petroleum ether, in a darkened room. Recrystallise it from water or EtOH. [Beilstein **17** V 239.]

**1,3-Dibromo-5,5-dimethylhydantoin** [77-48-5] **M 285.9, m 190-192°(dec), 190-193°(dec)**. Recrystallise it from H<sub>2</sub>O. Its solubility in CCl<sub>4</sub> is 0.003 mol/L at 25° and 0.024 mol/L at 76.5°. [Beilstein **24** III/IV 1101.]

**4',5'-Dibromofluorescein** [596-03-2] **M 490.1, m 285°**. Crystallise it from aqueous 30% EtOH. [Beilstein **19/6** V 462.]

**5,7-Dibromo-8-hydroxyquinoline** [521-74-4] **M 303.0, m 196°, pK<sub>1</sub><sup>25</sup> 5.84, pK<sub>2</sub><sup>25</sup> 9.56**. Crystallise it from acetone/EtOH. It can be sublimed. [Beilstein **21/3** V 290.]

**2,6-Dibromopyridine** [626-05-1] **M 236.9, m 117-119°, 118.5-119°, b 249°/757.5mm, pK<sub>Est</sub> <0**. Purify 2,6-dibromopyridine by steam distillation, then recrystallise it twice from EtOH. It does not form an HgCl<sub>2</sub> salt. [den Hertog & Wibaut *Rec Trav Chim Pays Bas* **51** 381 1932, Beilstein **20/5** V 435.]

**3,4-Dibromothiophene** [3141-26-2] **M 241.9, m 4-5°, b 94-95°/12mm, 212-213° (atm), 221-222°(atm), d<sub>4</sub><sup>25</sup> 2.188, n<sub>D</sub><sup>20</sup> 1.640**. Distil it in a vacuum, but if it is discoloured then dissolve it in Et<sub>2</sub>O, wash it with 2M NaOH solution, separate the layers, filter, dry the Et<sub>2</sub>O layer over CaCl<sub>2</sub>, filter, evaporate, and distil the residual oil in a vacuum. Nitration with Ac<sub>2</sub>O/HNO<sub>3</sub> at 60° yields *3,4-dibromo-2-nitrothiophene* **m 115-116°** (from EtOH). [Steinkopf et al. *Justus Liebigs Ann Chem* **512** 149-151 1934, Beilstein **17** III/IV 248, **17/1** V 308.]

**2,6-Di-tert-butyl-4-methylpyridine** [38222-83-2] **M 205.4, m 31-32°, 33-36°, b 148-153°/ 95mm, 223°/760mm, n<sub>D</sub><sup>20</sup> 1.476, pK<sub>Est</sub> ~5.7**. A possible impurity is 2,6-di-tert-butyl-4-neopentylpyridine. Attempts to remove coloured impurities directly by distillation, acid-base extraction or treatment with activated charcoal were unsuccessful. Pure material is obtained by dissolving 0.3mole of the alkyipyridine in pentane (150ml) and introducing it at the top of a cold water jacketed chromatographic column (40 x 4.5cm) (cooling is necessary because the base in pentane reacts exothermically with alumina) containing activated and acidic alumina (300g). The column is eluted with pentane using a 1L constant pressure funnel fitted at the top of the column to provide slight pressure. All the pyridine is obtained in the first two litres of eluent (the progress of elution is monitored by spotting a fluorescent TLC plate and examining under short wave UV light—a dark blue spot is evidence for the presence of the alkyipyridine). Elution is complete in 1 hour. Pentane is removed on a rotovap with 90-93% recovery yielding a liquid which solidifies on cooling, **m 31-32°**, and the base can be distilled. The *H<sub>2</sub>PtCl<sub>6</sub>* salt has **m 213-314°(dec)**, and the *CF<sub>3</sub>SO<sub>3</sub>H* salt has **m 202.5-203.5°** (from CH<sub>2</sub>Cl<sub>2</sub>). [Anderson & Stang *Org Synth Coll Vol VII* 144 1981, Beilstein **20/6** V 190.]

**2,6-Di-tert-butylpyridine** [585-48-8] **M 191.3, b 100-101°/23mm, d<sub>4</sub><sup>25</sup> 0.852, n<sub>D</sub><sup>20</sup> 1.473, pK<sub>25</sub> 5.02**. Redistil it from KOH pellets. [Beilstein **20** III/IV 2868.]

§ Polystyrene supported version is commercially available [107054-29-5] 200-400 mesh, crosslinked with 1% divinylbenzene containing ~1.8mmol N/g polymer.

**3,5-Dicarbethoxy-1,4-dihydro-2,4,6-collidine** [632-93-9] **M 267.3, m 131-132°**. Crystallise the ester from hot EtOH/water mixture. [Beilstein **22** H 147, **22** I 529, **22** II 100, **22** III/IV 1594.]

(+) and (-) (8,8-Dichlorocamphorylsulfonyl)oxaziridine [127184-05-8] M 298.2, m 178-180°, 183-186°,  $[\alpha]_D^{20}$  (+) and (-) 88.3° (c 1.3, CHCl<sub>3</sub>), (+) and (-) 91° (c 5, CHCl<sub>3</sub>). Recrystallise the enantiomers from EtOH [Davis & Weismiller *J Org Chem* 55 3715 1990].

**1,3-Dichloro-5,5'-dimethylhydantoin** [118-52-5] M 197.0, m 132-134°, 136°. Purify it by dissolving in conc H<sub>2</sub>SO<sub>4</sub> and diluting with ice H<sub>2</sub>O, collect the solid, dry it in a vacuum and recrystallise it from CHCl<sub>3</sub>. It sublimes at 100° in a vacuum. It exhibits time-dependent hydrolysis at pH 9. [Pettersson & Grzeskowiak *J Org Chem* 24 1414 1959, *Beilstein* 24 III/IV 1100.]

**4,5-Dichloro-3H-1,2-dithiol-3-one** [1192-52-5] M 187.1, m 52-56°, 61°, b 87°/0.5mm, 125°/11mm. Dissolve it in CH<sub>2</sub>Cl<sub>2</sub> (1g in 250ml), filter, wash it twice with H<sub>2</sub>O, evaporate and distil the residue *in vacuo* and then recrystallise it from petroleum ether. Its IR has  $\nu_{\max}$  at 1650 cm<sup>-1</sup>. [Boberg *Justus Liebigs Ann Chem* 679 109 1964, Boberg *Justus Liebigs Ann Chem* 693 212 1966, *Beilstein* 19/4 V 72.]

**5,7-Dichloro-8-hydroxyquinoline** [773-76-2] M 214.1, m 180-181°, pK<sub>1</sub> 1.89, pK<sub>2</sub> 7.62. Crystallise the dichloro-oxine from acetone/EtOH. [*Beilstein* 21 H 95.]

**6,9-Dichloro-2-methoxyacridine (3,9-Dichloro-7-methoxyacridine)** [86-38-4] M 278.1, m 160-161°, 164°, 163-165°. Crystallise it from \*benzene or 1,2-dichloroethane (m 162-163°). [Hall & Turner *J Chem Soc* 697 1945, *Beilstein* 21 III/IV 1553.]

**5,7-Dichloro-2-methyl-8-hydroxyquinoline (5,7-dichloro-8-hydroxyquinaldine)** [72-80-0] M 228.1, m 114-115°, pK<sub>Est(1)</sub> ~2.0, pK<sub>Est(2)</sub> ~8.4. Crystallise it from EtOH. [*Beilstein* 21/3 V 346.]

**4,6-Dichloro-5-nitropyrimidine** [4316-93-2] M 194.0, m 100-103°, 101-102°, pK<sub>Est</sub> <0. If too impure, then dissolve it in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dry it over MgSO<sub>4</sub>, evaporate it to dryness and recrystallise it from petroleum ether (b 85-105°) to give a light tan solid. It is soluble in *ca* 8 parts of MeOH [Boon et al. *J Chem Soc* 96 1951, Montgomery et al. in *Synthetic Procedures in Nucleic Acid Chemistry* Zorbach & Tipson eds, Wiley & Sons, NY, p76 1968]. [*Beilstein* 23 III/IV 899.]

**2,6-Dichloropurine** [5451-40-1] M 189.0, m 180-181.5°, 181°, 185-195°(dec), 188-189°, pK<sub>1</sub><sup>20</sup> 1.16 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>20</sup> 7.06. It can be recrystallised from 150 parts of boiling H<sub>2</sub>O and dried at 100° to constant weight. It is soluble in EtOAc. The HgCl<sub>2</sub> salt separates from EtOH solution. It has UV with  $\lambda_{\max}$  at 275nm ( $\epsilon$  8.9K) at pH 1; and 280nm ( $\epsilon$  8.5K) at pH 11 [Elion & Hitchings *J Am Chem Soc* 78 3508 1956, Schaeffer & Thomas *J Am Chem Soc* 80 3738 1958, Beaman & Robins *J Appl Chem (London)* 12 432 1962, Montgomery *J Am Chem Soc* 78 1928 1956]. [*Beilstein* 26 III/IV 1747.]

**2,6-Dichloropyridine** [2402-78-0] M 148.0, m 87-88°, pK<sup>25</sup> -2.86 (aqueous H<sub>2</sub>SO<sub>4</sub>). It crystallises from EtOH. [*Beilstein* 20/5 V 416.]

**3,5-Dichloropyridine** [2457-47-8] M 148.0, m 64-65°, 66-67°, b 178-178°/~760mm, pK<sup>25</sup> 0.67. Crystallise 3,5-dichloropyridine from EtOH, or distil it. [den Hertog et al. *Rec Trav Chim Pays Bas* 69 685 1950, *Beilstein* 20 H 23, 20 III/IV 2502, 20 V 417.]

**2,6-Dichloropyridine-3-carboxylic acid (2,6-dichloronicotinic acid)** [38496-18-3] M 192.0, m 146-148°, 147°, pK<sub>Est(1)</sub> ~-1. Purify the acid by recrystallisation from H<sub>2</sub>O or Et<sub>2</sub>O/petroleum ether (colourless needles), and the IR(KBr) has  $\nu_{\max}$  at 1738 cm<sup>-1</sup>, and  $\lambda_{\max}$  ( $\epsilon$ ) at 228.5(4700) and 273.5(8300)nm. The *acid chloride* [58584-83-1] M 210.4, b 72-74°/0.01mm, 117-118°/45mm, n<sub>D</sub><sup>20</sup> 1.5944, was obtained by treatment with SOCl<sub>2</sub>/DMF/\*C<sub>6</sub>H<sub>6</sub> and converted into the *amide* [70445-62-4] M 191.0, m 148-148.5°, or the *hydrazide* m 171-173° which crystallised from a large volume of MeCN. [Muterer et al. *Helv Chim Acta* 59 222 1976, Newcome et al. *J Org Chem* 44 2693 1979, *Beilstein* 22 H 44.] A pharmaceutical research building block acid of 90% grade is also available.

**2,6-Dichloropyridine-4-carboxylic acid (2,6-dichloroisonicotinic acid)** [5398-44-7] **M 192.0, m 210°, 209-212°, pK<sub>Est(1)</sub> ~ -1.** Purify the acid by crystallisation from H<sub>2</sub>O, and it forms needles or platelets from EtOH. The *methyl ester* crystallises from aqueous MeOH with **m 82°**, the *ethyl ester* from aqueous EtOH has **m 66°**, and the *amide* from H<sub>2</sub>O has **m 207-208°**. [Meyer et al. *Monatsh Chem* **36** 731 1915, Levelt & Wibaut *Rec Trav Chim Pays Bas* **48** 469 1929.]

**5,6-Dichloropyridine-3-carboxylic acid (5,6-dichloronicotinic acid)** [41667-95-2] **M 192.0, m 160-161°, 164-168°, 168°, pK<sub>Est(1)</sub> ~ 0.5.** Recrystallisation of the acid from H<sub>2</sub>O provides the *monohydrate* which becomes *anhydrous* (**m 161-162°**) on heating at 100°. The *acid chloride* distils at **125°/~24mm**, solidifies on cooling, and crystallises from petroleum ether with **m 48-49°**. The *methyl ester* has **m 67-68°** (from aqueous MeOH), and the *amide* forms colourless plates from aqueous EtOH with **m 218-220°**. [Räth & Schiffmann *Justus Liebigs Ann Chem* **487** 133 1931, Meyer & Graf *Chem Ber* **61** 2212 1928, *Beilstein* **22** H 44, **22** II 36, **22** III/IV 511, **22/2** V 80.]

**4,7-Dichloroquinoline** [86-98-6] **M 198.1, m 86.4-87.4°, b 148°/10mm, pK<sup>25</sup> 2.80.** Crystallise the dichloroquinoline from MeOH or 95% EtOH. [*Beilstein* **20/7** V 316.]

**2,3-Dichloroquinoxaline** [2213-63-0] **M 199.0, m 152-153°, 152-154°, pK<sub>Est</sub> < 0.** Recrystallise it from \*C<sub>6</sub>H<sub>6</sub> and dry it in a vacuum [Cheeseman *J Chem Soc* 1804 1955, *Beilstein* **23/7** V 144].

**cis-Dicyclohexyl-18-crown-6** [16069-36-6] **M 372.5, m 47-50°.** Purify it by chromatography on neutral alumina and elute with an ether/hexane mixture [see Izatt et al. *Inorg Chem* **14** 3132 1975]. Dissolve it in ether at *ca* 40°, and spectroscopic grade MeCN is added to the solution, which is then chilled. The crown ether precipitates and is filtered off. It is dried *in vacuo* at room temperature [Wallace *J Phys Chem* **89** 1357 1985]. [Vögtle ed. *Top Corr Chem* (Host Guest Complex Chemistry) **98** 1981.] **SKIN IRRITANT.**

**5,5-Diethylbarbituric acid (Barbital)** [57-44-3] **M 184.2, m 188-192°, pK<sub>1</sub><sup>25</sup> 8.02, pK<sub>2</sub><sup>25</sup> 12.7.** Crystallise barbital from water or EtOH and dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **24** III/IV 1901.]

**1,1'-Diethyl-2,2'-cyanine iodide** [977-96-8] **M 454.4, m 274°(dec).** It crystallises from EtOH and is dried in a vacuum oven at 80° for 4 hours. [*Beilstein* **23** II 267.]

**2,3-Dihydrobenzofuran (coumaran)** [496-16-2] **M 120.2, m -21.5°, 72-73°/12mm, 84°/17mm, 188°/atm, d<sub>4</sub><sup>20</sup> 1.065, n<sub>D</sub><sup>20</sup> 1.5524.** Suspend coumaran in aqueous NaOH and steam distil it. Saturate the distillate with NaCl and extract it with Et<sub>2</sub>O, dry the extract (MgSO<sub>4</sub>), filter, evaporate and distil the residue. It gives a strong violet colour with FeCl<sub>3</sub> + H<sub>2</sub>SO<sub>4</sub> and forms a yellow *picrate*, **m 76°** (from EtOH or \*C<sub>6</sub>H<sub>6</sub>) which loses coumaran in a desiccator [Bennett & Hafez *J Chem Soc* 287 1941, Baddeley et al. *J Chem Soc* 2455 1956]. [*Beilstein* **17/1** V 581.]

**Dihydropyran (3,4-dihydro-2H-pyran)** [110-87-2] **M 84.1, b 84.4°/742mm, 85.4-85.6°/760mm, d<sub>4</sub><sup>20</sup> 0.9261, n<sub>D</sub><sup>20</sup> 1.4423, pK<sub>Est</sub> ~ 4.2.** Partially dry dihydropyran with Na<sub>2</sub>CO<sub>3</sub>, then fractionally distil it. The fraction **b 84-85°** is refluxed with Na until hydrogen no longer evolves when fresh Na is added. It is then dried, and distilled again through a 60 x 1.2cm column packed with glass rings [Brandon et al. *J Am Chem Soc* **72** 2120 1950, UV: Elington et al. *J Chem Soc* 2873 1952, NMR: Bushweller & O'Neil *Tetrahedron Lett* 4713 1969]. It has been characterised as the *2-(3,5-dinitrobenzoyloxy)tetrahydropyrane* derivative, **m 103°** [prepared by dissolving 3,4-dinitrobenzoic acid (3g) in 50% excess of dihydropyran with warming, cooling, adding Et<sub>2</sub>O (5ml), whereby the derivative crystallises out quantitatively] which forms pale yellow crystals from 80% dihydropyran/Et<sub>2</sub>O [Woods & Kramer *J Am Chem Soc* **69** 2246 1947]. [*Beilstein* **17/1** V 181.]

**3,4-Dihydro-2H-pyrido[1,2a]-pyrimidin-2-one** [5439-14-5] **M 148.2, m 185-187°, 187-188°, 191-191.5°.** Dissolve it in CHCl<sub>3</sub>, filter, evaporate, then recrystallise the residue from EtOH/Me<sub>2</sub>CO (needles) which can be washed with Et<sub>2</sub>O and dried. It can also be recrystallised from CHCl<sub>3</sub>/petroleum ether or CHCl<sub>3</sub>/hexane. The *hydrochloride* has **m 295-295°** (dec, from EtOH or MeOH/Et<sub>2</sub>O), the *hydrobromide* has **m 299-300°**(dec) (from

MeOH/Et<sub>2</sub>O) and the *picrate* has **m** 224-226°(corr), **m** 219-220° from EtOH. [Adams & Pachter *J Am Chem Soc* **74** 4906 1952, Lappin *J Org Chem* **23** 1358 1958, Hurd & Hayao *J Am Chem Soc* **77** 115 1955, *Beilstein* **24** III/IV 299.]

**7,8-Dihydroxycoumarin (Daphnetin)** [486-35-1] **M 178.2, m 256°(dec), pK<sub>Est(1)</sub> ~8.5, pK<sub>Est(2)</sub> ~12.3.** Crystallise it from aqueous EtOH. It can be sublimed. [*Beilstein* **18/3** V 202.]

**trans-2,3-Dihydroxy-1,4-dioxane** [4845-50-5] **M 120.1, m 91-95°, 100°.** Recrystallise it from Me<sub>2</sub>CO. With phenylhydrazine it gives *glyoxal phenylhydrazone* **m** 175° (from Me<sub>2</sub>CO/petroleum ether). The *diacetyl* derivative has **m** 105-106° [Head *J Chem Soc* 1036 1955, Raudnitz *Chem Ind (London)* 166 1956]. [*Beilstein* **1** IV 3627.]

**2,5-Dihydroxy-1,4-dithiane** [40018-26-6] **M 152.2, m (142-147°) 150-152°, 151°.** Recrystallise the dithiane from EtOH. *2,5-Diethoxy-dithiane* has **m** 91° (92-93°); it crystallises from petroleum ether and can be sublimed at 60°/0.001mm [Hormatka & Haber *Monatsh Chem* **85** 1088 1954, Thiel et al. *Justus Liebigs Ann Chem* **611** 121 1958, Hesse & Jøeder *Chem Ber* **85** 924 1952]. [*Beilstein* **1** IV 3966.]

**S(-)-4',7-Dihydroxyflavanone (3,4'-dihydroxyisoflavone, Liquiritigenin)** [578-86-9] **M 256.3, m 203-205°** [ $\alpha$ ]<sub>D</sub><sup>20</sup> -225° (c 0.95, 95%EtOH). It crystallises from aqueous 50% EtOH. [*Beilstein* **18** III/IV 1780, **18/4** V 82.]

**5,7-Dihydroxy-4'-methoxyflavone (Acacetin)** [480-44-4] **M 284.3, m 261°, 260-265°.** Acacetin crystallises from 95% EtOH or acetic acid as pale yellow needles (m 263°). It is an anti-inflammatory. [Zemplen & Bogner *Chem Ber* **76** 452 1943, Pillon *Bull Soc Chim Fr* 9 1954, UV: Gaydou & Bianchini *Bull Soc Chim Fr* Part II 43 1978, *Beilstein* **18** III/IV 2683, **18/4** V 575.]

**(±)-7-(2,3-Dihydroxypropyl)theophylline (Diprophylline, Dyphylline)** [479-18-5] **M 254.3, m 158°, 160-164°, 161°, 161-164°, pK<sub>Est</sub> ~8.7.** Recrystallise it from EtOH or H<sub>2</sub>O. Its solubility in H<sub>2</sub>O is 33% at 25°, in EtOH it is 2% and in CHCl<sub>3</sub> it is 1%. Its UV has  $\lambda_{\max}$  (H<sub>2</sub>O) at 273nm ( $\epsilon$  8,855). [Roth *Arch Pharm* **292** 234 1959.] The *4-nitrobenzoyl* derivative has **m** 178° [Oshay *J Chem Soc* 3975 1956]. [*Beilstein* **26** III/IV 2370.]

**1,3-Diiminoisindoline** [3468-11-9] **M 145.2, m 193-194°(dec), 196°(dec), pK<sup>25</sup> 8.27.** It crystallises from H<sub>2</sub>O, MeOH or MeOH/Et<sub>2</sub>O (charcoal) in colourless prisms that become green on heating. [Elvidge & Linstead *J Chem Soc* 5000 1952]. Its IR (nujol) has bands at 3150 and 690 cm<sup>-1</sup>, and the UV has  $\lambda_{\max}$  at 251nm ( $\epsilon$  12,500), 256nm ( $\epsilon$  12,500) and 303nm ( $\epsilon$  4,600) [Elvidge & Golden *J Chem Soc* 700 1957, Clark et al. *J Chem Soc* 3593 1953]. The *thiocyanate* has **m** 250-255° (dec), the *monohydrochloride* has **m** 300-301° (turns green), and the *dihydrochloride* has **m** 326-328° (turns green) and the *picrate* crystallises from EtOH with **m** 299° (dec). [*Beilstein* **22/13** V 5.]

**5,7-Diiodo-8-hydroxyquinoline** [83-73-8] **M 397.0, m 214-215°(dec) pK<sub>Est(1)</sub> ~3.2, pK<sub>Est(2)</sub> ~8.2.** It crystallises from xylene and is dried at 70° in a vacuum. [*Beilstein* **21** II 58.]

**6-Dimethylaminopurine** [938-55-6] **M 163.1, m 257.5-258.5°, 259-262°, 263-264°, pK<sub>1</sub><sup>25</sup> 3.87, pK<sub>2</sub><sup>25</sup> 10.5.** It is purified by recrystallisation from H<sub>2</sub>O, EtOH (0.32g in 10ml) or CHCl<sub>3</sub>. [Albert & Brown *J Chem Soc* 2060 1954, UV: Mason *J Chem Soc* 2071 1954.] The *monohydrochloride* crystallises from EtOH/Et<sub>2</sub>O, **m** 253°(dec) [Elion et al. *J Am Chem Soc* **74** 411 1952], the *dihydrochloride* has **m** 225°(dec) and the *picrate* has **m** 245° (235-236.5°) [Fryth et al. *J Am Chem Soc* **80** 2736 1958]. [*Beilstein* **26** III/IV 3566.]

**1,3-Dimethylbarbituric acid** [769-42-6] **M 156.1, m 123°, pK<sup>25</sup> 4.56.** Crystallise the acid from water and sublime it in a vacuum. Also purify it by dissolving 10g in 100ml of boiling CCl<sub>4</sub>/CHCl<sub>3</sub> (8:2) (1g charcoal), filtering and cooling to 25°. Dry it *in vacuo* [Kohn et al. *Anal Chem* **58** 3184 1986]. [*Beilstein* **24** III/IV 1875.]

**5,6-Dimethylbenzimidazole** [582-60-5] **M 146.2, m 205-206°, pK<sub>1</sub><sup>25</sup> 5.96, pK<sub>2</sub><sup>25</sup> 12.52.** Crystallise 5,6-dimethylbenzimidazole from diethyl ether. It sublimes at 140°/3mm. [*Beilstein* **23/6** V 454.]

**2,3-Dimethylbenzothiophene** [31317-17-6] M 212.3, m 69-70°, b 123-124°/10mm, n<sub>D</sub><sup>19</sup> 1.6171. Fractionate it through a 90cm Monel spiral column, or other efficient fractionating or spinning band column and collect the middle fraction. It has also been purified by chromatography on basic alumina using pentane as eluent. [Tedjanulia et al. *J Heterocycl Chem* **20** 1485 1983, *Beilstein* **17** III/IV 161.]

**4,4'-Dimethyl-2,2'-bipyridine** [1134-35-6] M 184.2, m 175-176°, pK<sub>Est(1)</sub> ~0.2, pK<sub>Est(2)</sub> ~4.9. Crystallise it from ethyl acetate. [Elliott et al. *J Am Chem Soc* **107** 4647 1985, *Beilstein* **23/8** V 79.]

**1,1'-Dimethyl-4,4'-bipyridylum dichloride (3H<sub>2</sub>O; Methyl Viologen Dichloride, paraquat dichloride)** [1910-42-5] M 311.2, m >300°(dec). Recrystallise the dichloride from MeOH/acetone mixture. It has also been recrystallised three times from absolute EtOH [Bancroft et al. *Anal Chem* **53** 1390 1981]. Dry it at 80° in a vacuum. [*Beilstein* **23/8** V 30.]

**1,3-Dimethylbutadiene sulfone (1,3-dimethylsulfolene, 2,5-dihydro-2,4-dimethylthiophene)** [10033-92-8] M 145.2, m 40.4-41.0°. Crystallise it from Et<sub>2</sub>O, Et<sub>2</sub>O/pentane or CCl<sub>4</sub>. [Grummitt et al. *J Am Chem Soc* **72** 5167 1950, Bartlett et al. *J Org Chem* **32** 1290 1967, *Beilstein* **17** III/IV 161.]

**2,2-Dimethyl-1,3-dioxan-4,6-dione (Meldrum's Acid)** [2033-24-1] M 144.1, m 92-96°, 94-95°, pK<sup>25</sup> 5.1, 7.32. Crystallise the dione from Me<sub>2</sub>CO/H<sub>2</sub>O. It is a useful synthon for the C3 malonic acid moiety. [Arnette et al. *J Am Chem Soc* **106** 6759 1984, Bihlmayer et al. *Monatsh Chem* **98** 564 1967, Review: McNab *Chem Soc, Rev* **7** 345 1978, Chan & Huang *Synthesis* 452 1982, *Beilstein* **19/5** V 8.] It is synthesised by a modified procedure in which concentrated H<sub>2</sub>SO<sub>4</sub> (1.5ml) is added to a stirred suspension of powdered malonic acid (52g, 0.5 mole) in acetic anhydride (60ml, 0.6 mole), whereby the malonic acid dissolves with spontaneous cooling. Acetone (40ml, 0.55 mole) is added to the mixture while keeping the temperature at 20-25°, then cooled in a refrigerator overnight, the crystals are filtered off, washed three times with ice-water (enough to cover the crystals) and air dried. The crude product (35g, 49%) is recrystallised without heating by dissolving it (10g) in Me<sub>2</sub>CO (20ml), filtering, and adding H<sub>2</sub>O (40ml), with ~70% recovery of material m 94-95°. [Pihlaja & Seilo *Acta Chem Scand* **22** 3053 1968.]

*Alternatively* concentrated H<sub>2</sub>SO<sub>4</sub> (0.5ml) is added dropwise to a stirred suspension of powdered malonic acid (52g, 0.5 mole) in redistilled isoprenyl acetate (62ml, 55g, [108-22-5]), when the temperature rose from 23° to 31° in 45 minutes and all the solid dissolved within 1 hour. Treatment of the reaction as above gave Meldrum's acid (37g, 50%). [Davidson & Bernhard *J Am Chem Soc* **70** 3426 1948]. Its <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has δ at 1.73 (s, 2-Me<sub>2</sub>) and 3.60 (s, 5-H) [Schuster & Schuster *Tetrahedron* **25** 199 1969].

**3,6-Dimethyl-1,4-dioxan-2,5-dione (Lactide)** [*cis-RS,RS*(±) 615-95-2, *cis-R,R*(+) 95-96-5, *cis-S,S*(-) 4511-42-6] M 144.1. This is the cyclic dilactone of lactic acid. The (±)-*cis-racemate* has been distilled with b 142°/8mm; the distillate which solidifies gives yellow needles on recrystallisation from EtOH with m 128°, from Et<sub>2</sub>O with m 129°, or CHCl<sub>3</sub> with m 126°, with IR ν<sub>max</sub> at 1720-1740 cm<sup>-1</sup>. It hydrolyses in cold H<sub>2</sub>O. [Carothers et al. *J Am Chem Soc* **54** 772 1932]. A *trans-form* (probably *RS,SR*) has been reported which crystallises from Et<sub>2</sub>O with m 42-43° [Hummel et al. *Acta Cryst (Sect B)* **38** 1679 1982]. The *R,R*(+)-*lactide* has b 150°/2mm and crystallises from Et<sub>2</sub>O with m 95°, or m 96.5-97.5° (from CHCl<sub>3</sub>) or m 97.7° (from EtOAc) and [α]<sub>D</sub><sup>22</sup> +297° (c 1.2, \*C<sub>6</sub>H<sub>6</sub>). The *S,S*(-)-*lactide* has b 150°/2.5mm and crystallises from EtOAc with m 98.7°, or m 95° (from CHCl<sub>3</sub>) or m 96.5-97.5° (from CCl<sub>4</sub>) and [α]<sub>D</sub><sup>22</sup> -297° (c 1.2, \*C<sub>6</sub>H<sub>6</sub>). [Toniolo et al. *J Org Chem* **35** 6 1970, *Beilstein* **19** H 154, **19** I 179, **19** II 176, **19** IV 1927, **19/5** V 10.]

**2,9-Dimethyl-4,7-diphenyl-1,10-phenanthroline (Bathocuproine)** [4733-39-5] M 360.5, m >280°, pK<sub>Est</sub> ~5.6. Purify it by recrystallisation from \*benzene. [Smith & Wilkins *Anal Chem* **25** 510 1953, *Beilstein* **23** III/IV 2160.]

**2,2-Dimethylethyleneimine (2,2-dimethylaziridine)** [2658-24-4] M 71.1, b 70.5-71.0°, 72°, n<sub>D</sub><sup>20</sup> 1.405, pK<sup>25</sup> 8.64. Dry the 2,2-dimethylaziridine over solid KOH, filter and freshly distil from sodium before use, and store it under dry nitrogen. The *N*-phenylthiocarbonyl derivative crystallises from petroleum ether containing a trace of Me<sub>2</sub>CO with m 92.5-93.5°. [Hassner et al. *J Am Chem Soc* **91** 5046 1969, Cairns *J Am Chem Soc* **63** 871 1941, Lamaty et al. *Justus Liebigs Ann Chem* **726** 77 1969, *Beilstein* **20** III/IV 280.]

**5,5-Dimethylhydantoin** [77-71-4] M 128.1, m 177-178° pK<sup>24</sup> 9.19. Crystallise the hydantoin from EtOH and sublime it *in vacuo*. [Beilstein 24 III/IV 1097.]

**4,6-Dimethyl-2-hydroxypyrimidine** [108-79-2] M 124.1, m 198-199°, 202-205°, pK<sub>1</sub><sup>20</sup> 3.77, pK<sub>2</sub><sup>20</sup> 10.50. Crystallise the pyrimidine from absolute EtOH (charcoal). [Beilstein 24/2 V 138.]

**1,2-Dimethyl-1H-imidazole** [1739-84-0] M 96.1, m 38-40°, b 206°/760mm, d<sub>4</sub><sup>20</sup> 1.084, pK<sub>Est</sub> ~8.1. Crystallise the imidazole from \*benzene, dry and store it at 0-4°. The *picrate* crystallises from H<sub>2</sub>O or EtOH with m 181°. [Balaban & Pymann *J Chem Soc* 125 1570 1924, Gorun et al. *J Am Chem Soc* 109 4244 1987, Beilstein 23 H 66, 23 II 56, 23 III/IV 594.]

**1,3-Dimethyl-2-imidazolinone (DMI, N,N'-dimethylethyleneurea, DMEU)** [80-73-9] M 114.2, m 8.2°, b 67-68°/2mm, 104°/5mm, 106-107°/17mm, 224-226°/atm, d<sub>4</sub><sup>25</sup> 1.056, n<sub>D</sub><sup>20</sup> 1.472. After preparation by reaction of N,N'-dimethylethylenediamine in toluene with phosgene in toluene below 15°, excess phosgene is removed by blowing air through the mixture. The hydrochloride salt that separates is removed by filtration, and washed with CHCl<sub>3</sub>. The combined filtrate and washings are evaporated off, and the residue is distilled (b 129°/39mm). The crude distillate is treated with H<sub>2</sub>O; K<sub>2</sub>CO<sub>3</sub> is added to saturate the solution, and extracted with CHCl<sub>3</sub>, the extract is dried (K<sub>2</sub>CO<sub>3</sub>), filtered, and pure 1,3-dimethylimidazolin-2-one distils at b 104°/5mm. It has been used as an alternative to HMPA [see 680-31-9] as a high dielectric solvent for reactions. [Boon *J Chem Soc* 307 1947, Lien & Kumler *J Med Chem* 11 241 1968, Kohn et al. *J Org Chem* 42 941 1977, Beilstein 24 III/IV 9.]

**3,3-Dimethyloxetane** [6921-35-3] M 86.1, b 79.2-80.3°/760mm, d<sub>4</sub><sup>20</sup> 0.836, n<sub>D</sub><sup>20</sup> 1.399. Purify 3,3-dimethyloxetane by gas chromatography using a 2m silicone oil column or distil it. Fractionate it at atmospheric pressure (preferably under N<sub>2</sub> or Ar). [Beilstein 17 II 21.]

**2,9-Dimethyl-1,10-phenanthroline (neocuproine hemihydrate)** [484-11-7] M 208.3, 217.3 (hemihydrate), m 162-164°, pK<sup>25</sup> 5.85. Purify it as the *hemihydrate* by crystallisation from water and as the *anhydrous* base from \*benzene. [Beilstein 23/8 V 527.]

**4,4-Dimethyl-2,6-piperidinedione (4,4-dimethylglutarimide)** [1123-40-6] M 141.2, m 144-146°, pK<sub>Est</sub> ~11.5. Recrystallise the imide from hot H<sub>2</sub>O or EtOH [Arnett & Harrelson *J Am Chem Soc* 109 809 1987]. [Beilstein 21 H 391, 21 I 331, 21 II 309, 21 III/IV 4601, 21/9 V 592.]

**2,5-Dimethylpyrazine** [123-32-0] M 108.1, b 156°, d<sub>4</sub><sup>20</sup> 0.990, n<sub>D</sub><sup>20</sup> 1.502, pK<sub>1</sub><sup>25</sup> -4.6 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sub>2</sub><sup>25</sup> 1.85. Purify it *via its picrate* (m 150°) which is decomposed with a base (e.g. KOH) and distilled. [Wiggins and Wise *J Chem Soc* 4780 1956]. [Beilstein 23/5 V 403.]

**3,5-Dimethylpyrazole** [67-51-6] M 96.1, m 107-108°, pK<sup>20</sup> 4.16. Crystallise it from cyclohexane or water. [Barszcz et al. *J Chem Soc, Dalton Trans* 2025 1986, Beilstein 23/5 V 110.]

**2,3-Dimethylquinoxaline** [2379-55-7] M 158.2, m 106°, pK<sup>25</sup> -3.84 (aqueous H<sub>2</sub>SO<sub>4</sub>). It has been purified by steam distillation with the base crystallising in the distillate. Recrystallise it from distilled water or aqueous EtOH. The *sulfate* crystallises from EtOH with m 151-152°(dec). [Gibson *J Chem Soc* 343 1927, Beilstein 23 H 191, 23 II 197, 23 III/IV 1277.]

**2,4-Dimethylsulfolane** [1003-78-7] M 148.2, b 128°/77mm, d<sub>D</sub><sup>25</sup> 1.1314, n<sub>D</sub><sup>20</sup> 1.474. Distil the 2,4-dimethylsulfolane in a vacuum. [Beilstein 17/1 V 82.]

**1,3-Dimethyluracil** [874-14-6] M 140.1, m 121-122°, pK<sup>25</sup> -3.25 (aqueous H<sub>2</sub>SO<sub>4</sub>). Crystallise it from EtOH/ether. [Beilstein 24 III/IV 1196.]

**9,9-Dimethylxanthen (4,4-dimethyl-2,3:5,6-dibenzopyran)** [19814-75-6] M 210.3, m 36-38°, b 114-115°/0.6mm, n<sub>D</sub><sup>20</sup> 1.5973. It is prepared under argon (Schlenk equipment) by adding Me<sub>3</sub>Al solution (2.0M in

toluene, 320ml, 0.64mmol), *via* a septum over 50 minutes, to an ice-water cooled stirred suspension of 9-xanthone (50.0g, 0.255mol) in toluene (300ml) (exothermic reaction, no gas evolved). The dark red solution is allowed to warm to  $\sim 25^\circ$  over 3 hours, and stirred for a further 14 hours. (*Alternatively* it can be heated to  $60^\circ$  for  $\sim 0.5$  hours.) The mixture is transferred (*via* a cannula under argon pressure) to a stirred mixture of concentrated HCl (250ml) and ice (4 L). The organic phase is separated, dried ( $\text{MgSO}_4$ ), filtered, the solvent is evaporated *in vacuo* to give a yellow oil (51.5g, 96%) which distils at  $114\text{--}115^\circ/0.6\text{mm}$ , and solidifies on cooling. The  $^1\text{H NMR}$  ( $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.63 (s, 6H,  $\text{CH}_3$ ), 7.0-7.6 (m, 8H, aromatic-H); MS has  $m/z$  at 210 ( $\text{M}^+$ ), 195 (base peak,  $\text{M}^+ - \text{CH}_3$ ). [Nowick et al. *J Am Chem Soc* **112** 8904 1990, Meisters & Mole *Aust J Chem* **27** 1655 1974, *Beilstein* **17** II 287.]

**1,3-Dioxolan-2-one (ethylene carbonate)** [96-49-1] **M 88.1, m  $37^\circ$ ,  $39^\circ$ ,  $40^\circ$ , b  $65\text{--}67^\circ/1\text{mm}$ ,  $126^\circ/17\text{mm}$ ,  $238^\circ/760\text{mm}$ ,  $d_4^{20}$  1.321,  $n_D^{40}$  1.4199.** Dry 1,3-dioxolan-2-one over  $\text{P}_2\text{O}_5$ , then fractionally distil it at low or atmospheric pressure. Recrystallise it from dry  $\text{Et}_2\text{O}$  (plates, **m  $36.5^\circ$ ,  $38.5\text{--}40^\circ$**  was also reported). It is soluble in  $\text{H}_2\text{O}$ . [*Beilstein* **19** II 135, **19 III/IV** 1556, **19/4 V** 6.]

**1,3-Dioxane** [505-22-6] **M 88.1, b  $104.5^\circ/751\text{mm}$ ,  $d_4^{20}$  1.040,  $n_D^{20}$  1.417.** Dry the dioxane with sodium and fractionally distil it. [*Beilstein* **19/1 V** 11.]

**1,4-Dioxane** [123-91-1] **M 88.1, f  $11.8^\circ$ , b  $101.3^\circ$ ,  $d_4^{25}$  1.0292,  $n_D^{15}$  1.4236,  $n_D^{25}$  1.42025.** It is prepared commercially either by dehydration of ethylene glycol with  $\text{H}_2\text{SO}_4$  and heating ethylene oxide or bis( $\beta$ -chloroethyl)ether with NaOH. The usual impurities are acetaldehyde, ethylene acetal, acetic acid, water and peroxides. Peroxides can be removed (and the aldehyde content decreased) by percolation through a column of activated alumina (80g per 100-200ml solvent), by refluxing with  $\text{NaBH}_4$  or anhydrous stannous chloride and distilling, or by acidification with conc HCl, shaking with ferrous sulfate and leaving in contact with it for 24 hours before filtering and purifying further.

Hess and Frahm [*Chem Ber* **71** 2627 1938] refluxed 2L of dioxane with 27ml conc HCl and 200ml water for 12 hours with slow passage of nitrogen to remove acetaldehyde. After cooling the solution, KOH pellets were added slowly and with shaking until no more would dissolve and a second layer had separated. The dioxane was decanted, treated with fresh KOH pellets to remove any aqueous phase, then transferred to a clean flask where it was refluxed for 6-12 hours with sodium, then distilled from it. *Alternatively*, Kraus and Vingee [*J Am Chem Soc* **56** 511 1934] heated it on a steam bath with solid KOH until fresh addition of KOH gave no more resin (due to acetaldehyde). After filtering through paper, the dioxane was refluxed over sodium until the surface of the metal was not further discoloured during several hours. It was then distilled from sodium.

The acetal (**b  $82.5^\circ$** ) is removed during fractional distillation. Traces of \*benzene, if present, can be removed as the \*benzene/MeOH azeotrope by distillation in the presence of MeOH. Distillation from  $\text{LiAlH}_4$  removes aldehydes, peroxides and water. Dioxane can be dried using Linde type 4X molecular sieves. Other purification procedures include distillation from excess  $\text{C}_2\text{H}_5\text{MgBr}$ , refluxing with  $\text{PbO}_2$  to remove peroxides, fractional crystallisation by partial freezing and the addition of KI to dioxane acidified with aqueous HCl. Dioxane should be stored out of contact with air, preferably under  $\text{N}_2$ .

*A detailed purification procedure is as follows:* Dioxane is stood over ferrous sulfate for at least 2 days, under nitrogen. Then water (100ml) and conc HCl (14ml)/ litre of dioxane are added (giving a pale yellow colour). After refluxing for 8-12 hours with vigorous  $\text{N}_2$  bubbling, pellets of KOH are added to the warm solution to form two layers and to discharge the colour. The solution is cooled rapidly with more KOH pellets being added (magnetic stirring) until no more dissolved in the cooled solution. After 4-12 hours, if the lower phase is not black, the upper phase is decanted rapidly into a clean flask containing sodium, and refluxed over sodium (until freshly added sodium remained bright) for 1 hour. The middle fraction is collected (and checked for minimum absorbency below 250nm). The distillate is fractionally frozen three times by cooling in a refrigerator, with occasional shaking or stirring. This material is stored in a refrigerator. Before use it is thawed, refluxed over sodium for 48 hours, and distilled into a container. All joints are clad with Teflon tape.

Coetzee and Chang [*Pure Appl Chem* **57** 633 1985] dried the solvent by passing it slowly through a column (20g/L) of 3A molecular sieves activated by heating at  $250^\circ$  for 24 hours. Impurities (including peroxides) are removed by passing the effluent slowly through a column packed with type NaX zeolite (pellets ground to

0.1mm size) activated by heating at 400° for 24 hours or chromatographic grade basic Al<sub>2</sub>O<sub>3</sub> activated by heating at 250° for 24 hours. After removal of peroxides the effluent is refluxed for several hours over sodium wire, excluding moisture, distilled under nitrogen or argon and stored in the dark.

One of the best tests of purity of dioxane is the formation of the purple disodium benzophenone complex during reflux and its persistence on cooling. (Benzophenone is better than fluorenone for this purpose and for the storing of the solvent.) [Carter et al. *Trans Faraday Soc* **56** 343 1960, *Beilstein* **19** V 16.] **TOXIC.**

**Rapid purification:** Check for peroxides (see Chapter 1 and Chapter 2 for test under ethers). Pre-dry with CaCl<sub>2</sub> or better over Na wire. Then reflux the pre-dried solvent over Na (1% w/v) and benzophenone (0.2% w/v) under an inert atmosphere until the blue colour of the benzophenone ketyl radical anion persists. Distil. and store it over 4A molecular sieves in the dark.

**1,3-Dioxolane** [646-06-0] **M** 74.1, **b** 75.0-75.2°, **d**<sub>4</sub><sup>20</sup> 1.0600, **n**<sub>D</sub><sup>21</sup> 1.3997. Dry it with solid NaOH, KOH or CaSO<sub>4</sub>, and distil it from sodium or sodium amalgam. Barker et al. [*J Chem Soc* 802 1959] heated 34ml of dioxolane under reflux with 3g of PbO<sub>2</sub> for 2 hours, then cooled and filtered. After adding xylene (40ml) and PbO<sub>2</sub> (2g) to the filtrate, the mixture is fractionally distilled. Addition of xylene (20ml) and sodium wire to the main fraction (**b** 70-71°) led to a vigorous reaction, following which the mixture was again fractionally distilled. Xylene and sodium additions are made to the main fraction (**b** 73-74°) before it is finally distilled. [*Beilstein* **19**/I V 6.]

**5,5-Diphenylhydantoin** [57-41-0] **M** 252.3, **m** 293-295°. Crystallise the hydantoin from EtOH. [*Beilstein* **24** III/IV 1748.]

**1,3-Diphenylisobenzofuran** [5471-63-6] **M** 270.3, **m** 129-130°. Recrystallise it from EtOH or EtOH CHCl<sub>3</sub> (1:1) under red light (as in photographic dark rooms) or from \*benzene in the dark. [*Beilstein* **17**/2 V 503.]

**2,5-Diphenyl-1,3,4-oxadiazole (PPD)** [725-12-2] **M** 222.3, **m** 70° (hydrate), 139-140° (anhydrous), **b** 231°/13mm, 248°/16mm. Its solubility in CHCl<sub>3</sub> is 10%. Crystallise it from EtOH and sublime it *in vacuo*. [*Beilstein* **27** III/IV 2712.]

**2,5-Diphenyloxazole (PPO)** [92-71-7] **M** 221.3, **m** 74°, **b** 360°/760mm. Distil it in steam and crystallise it from ligroin. [*Beilstein* **27** III/IV 1437.]

**2RS,5RS-(±)-2,5-Diphenylpyrrolidine (± trans isomer)** [22147-84-8] **M** 223.3, **b** 136°/0.3mm, **n**<sub>D</sub><sup>25</sup> 1.5866, **pK**<sub>Est</sub> ~8.0. A mixture of *cis*-(*meso*) and *trans*-(±) pyrrolidines (**b** ~130-145°/0.15-0.2mm, ~138g, 0.6mol) in EtOH (210ml) and concentrated HCl (100ml) was treated dropwise, during ~2 hours, with a solution of NaNO<sub>2</sub> (61.5g, 0.9mol) in H<sub>2</sub>O (100ml) at ~0° (ice-bath cooling). The mixture was poured into H<sub>2</sub>O (1L) and the precipitate (~160g) was collected, washed with H<sub>2</sub>O and fractionally recrystallised from Me<sub>2</sub>CO to give the *trans*-nitroso derivative (**m** 140-140.9°, 51%) [22147-82-6], and the *cis*-nitroso derivative (**m** 97-98°, 20%) [22147-81-5]. This *trans*-nitroso-pyrrolidine was de-nitrosated (HCl gas in Et<sub>2</sub>O) to the *trans*-hydrochloride **m** 196.5-180° which was recrystallised from EtOH/Et<sub>2</sub>O. The *trans*-free base was obtained by basifying an aqueous solution to pH 9-10 and extracting with Et<sub>2</sub>O (3 x 50ml), drying the extract (Mg<sub>2</sub>SO<sub>4</sub>), filtering, evaporating and distilling the residual oil (**b** 136°/0.3mm). The IR (film) has **v**<sub>max</sub> at 1110-1025 (three peaks, one is for C-N) and 3330 (N-H) cm<sup>-1</sup>. Similarly the *cis* (*meso*) *N*-nitroso derivative gave the *2RS,5SR*-diphenylpyrrolidine *meso*-hydrochloride which had **m** 224.7-225.5°, and the *meso*-free base [22147-83-7] had **b** 121-123.4°/0.2mm, **n**<sub>D</sub><sup>25</sup> 1.5850. The optically active *2R,5R*-(*-*)-diphenylpyrrolidine enantiomer [155155-73-0] had **m** 49-53°, [**α**]<sub>D</sub><sup>22</sup> +118° (c 1.0, CHCl<sub>3</sub>). [Overberger et al. *J Am Chem Soc* **91** 687 1969. Breuer & Melamed *J Org Chem* **37** 3949 1972.]

(±)-**α, α-Diphenyl-2-pyrrolidinemethanol (α, α-diphenyl-DL-prolinol)** [63401-04-7, 112022-88-55] **M** 253.3, **m** 82-83° (76° also reported), **pK**<sub>Est</sub> ~9.5. It is prepared as the (*±*)-hydrochloride [16226-54-3] **M** 289.8, **m** 267-269° (dec, 262-263° was also reported) by dissolving the amine in Et<sub>2</sub>O. Et<sub>2</sub>O MeOH or CH<sub>2</sub>Cl<sub>2</sub>, and dry HCl gas is bubbled through (avoid excess HCl as it may dissolve precipitated NH<sub>4</sub>Cl), filtering,

evaporating, and the residue is recrystallised from MeOH/Me<sub>2</sub>CO; and from which the *free base* is obtained by treatment with aqueous NaOH and extraction with Et<sub>2</sub>O. It has psychostimulating activity. [Enders et al. *Org Synth* **58** 113 1978, Enders et al. *Org Synth Coll Vol VI* 542 1988, Enders et al. *Synthesis* 548 1976, Likhosherstov et al. *Khim Farm Z* **1** 30 1967, *Chem Abstr* **67** 90642 1967, *Beilstein* **21** III IV 1519.]

*R*-(+)- $\alpha$ ,  $\alpha$ -Diphenyl-2-pyrrolidinemethanol ( $\alpha$ ,  $\alpha$ -Diphenyl-D-prolinol) [22348-32-0] **M253.3**, *m* 78-80°,  $[\alpha]_D^{20}$  +69° (c 3, CHCl<sub>3</sub>). It is purified by recrystallisation from EtOH or hexane. It was also purified via the (*R*)-hydrochloride [172152-19-1] as above.

*S*-(-)- $\alpha$ ,  $\alpha$ -Diphenyl-2-pyrrolidinemethanol ( $\alpha$ ,  $\alpha$ -diphenyl-L-prolinol) [112108-01-9] **M253.3**, *m* 79-9.5°, 83°,  $[\alpha]_D^{20}$  -68.1° (c 3.2, CHCl<sub>3</sub>),  $[\alpha]_D^{20}$  -87.5° (c 1.16, CH<sub>2</sub>Cl<sub>2</sub>). It is purified by recrystallisation from EtOH or hexane. The *S*-(-)-hydrochloride [148719-90-8] melts above 240°, and the *benzoyl* derivative *m* 183° crystallises from H<sub>2</sub>O, MeOH or EtOH. [Enders et al. *Bull Soc Chem Belg* **97** 691 1988, Kerrick & Beak *J Am Chem Soc* **113** 9708 1991, Corey & Bakshi *Tetrahedron Lett* **31** 611 1990, Mathre *J Org Chem* **58** 2880 1993, *Beilstein* **21** III IV 1519.]

Dipicolinic acid (pyridine-2,6-dicarboxylic acid) [490-83-2] **M 167.1**, *m* 255°(dec),  $\lambda_{\max}$  270nm,  $pK_1^{20}$  2.10,  $pK_2^{20}$  4.68. Crystallise the acid from water, and sublime it in a vacuum. [*Beilstein* **22/4 V** 128.]

Di-(4-pyridyl)hydrazine (*N,N'*-disonicolinoyl hydrazine) [4320-75-3] **M 246.2**, *m* 254-255°, 259-260°. Crystallise it from water, aqueous EtOH or propan-1-ol. [Albert & Rees *Biochem J* **61** 128 1955, *Beilstein* **22** III IV 663.]

2,2'-Dipyridylamine [1202-34-2] **M 171.2**, *m* 84° and remelts at 95° after solidifying, *b* 176-178°/13mm, 222°/50mm, 307-308°/60mm,  $pK^{25}$  6.69 (in 20% aqueous EtOH). Crystallise the amine from \*benzene or toluene [Blakley & De Armond *J Am Chem Soc* **109** 4895 1987]. The amine is also recrystallised from Me<sub>2</sub>CO (*m* 95.1°) or distilled in a vacuum. [*Beilstein* **22** I 630, **22** II 331, **22** III IV 3961.]

2,2'-Dipyridyl disulfide (2,2'-dithiopyridine, Aldrithiol-2) [2127-03-9] **M 220.3**, *m* 53°, 57-58°,  $pK_1^{25}$  0.35,  $pK_2^{25}$  2.45. Recrystallise the disulfide H<sub>2</sub>O from \*C<sub>6</sub>H<sub>6</sub>, petroleum ether (6:7), ligroin or \*C<sub>6</sub>H<sub>6</sub>. The *picrate* has *m* 119° (from EtOH). [Walter et al. *Justus Liebigs Ann Chem* **695** 7785 1966, Marckwald et al. *Chem Ber* **33** 1556 1900, Brocklehurst & Little *Biochem J* **133** 67.78 1973, *Beilstein* **21** III IV 48.] It has been used as a 1mM solution in EtOH for the spectrophotometric estimation of thiols. Essentially the thiol displaces half the disulfide molecule liberating the 2-mercaptopyridine anion, thereby shifting the  $\lambda_{\max}$  from 340nm (of the disulfide) to 268nm (of the anion) at pH 9, or 278nm in H<sub>2</sub>O. (Compare with 4,4'-dipyridyl disulfide (below) which has been used for the same purpose [Humphrey et al. *Anal Chem* **40** 698 1970].

4,4'-Dipyridyl disulfide (4,4'-dithiopyridine, Aldrithiol-4) [2645-22-9] **M 220.3**, *m* 74-76°,  $pK_{\text{Est}(1)}$  ~1.5,  $pK_{\text{Est}(2)}$  ~4.5. Recrystallise the disulfide from H<sub>2</sub>O, EtOH, Me<sub>2</sub>CO, \*C<sub>6</sub>H<sub>6</sub> or petroleum ether. It has been used as a 1mM solution in EtOH for the spectrophotometric estimation of thiols. Essentially the thiol displaces half the disulfide molecule liberating the 4-mercaptopyridine (4-pyridinethiol) anion, thereby shifting the  $\lambda_{\max}$  from 324nm (of the disulfide) to 285nm (of the anion) at pH 9. (Compare with 2,2' dipyridyl disulfide above which has been used for the same purpose.) [Humphrey et al. *Anal Chem* **40** 698 1970, Cheng & Ritchie *Aust J Chem* **26** 1785 1973.] [*Beilstein* **21** II 35, for review see Aldrich in *Aldrichimica Acta* **4** 33 1971.]

1,2-Di-(4-pyridyl)-ethane [4916-57-8] **M 184.2**, *m* 110.9-111.2°, 114.5-116°, *b* 167-174°/3mm,  $pK_{\text{Est}(1)}$  ~3.8,  $pK_{\text{Est}(2)}$  ~5.4. Crystallise the ethane from cyclohexane/\*benzene (3:1, solubility is ~7.5g/100ml). The *dihydrochloride* crystallises from EtOH with *m* 329-330°(dec). [Bergmann et al. *J Am Chem Soc* **74** 5981 1952, Thayer & Carson *J Am Chem Soc* **70** 2331 1948, Jampolsky et al. *J Am Chem Soc* **74** 5222 1952, Chow & Fouss *J Am Chem Soc* **80** 1095 1958, *Beilstein* **23** III IV 1389.]

*trans*-1,2-Di-(4-pyridyl)-ethylene [13362-78-2] **M 182.2**, *m* 153-154°, 155.5-156.5°,  $pK_1^{25}$  3.65,  $pK_2^{25}$  5.6. Crystallise the ethylene from water (1.6g/100ml at 100°). The *di-hydrochloride* has *m* 347°, from EtOH.

[*Beilstein* 23/8 V 239.]

**1,3-Di-(4-pyridyl)-propane** [17252-5]-6] *M* 198.3, *m* 60.5-61.5°, 62-65°,  $pK_{\text{Est}(1)} \sim 4.5$ ,  $pK_{\text{Est}(2)} \sim 5.5$ . Crystallise the propane from *n*-hexane \*benzene (5:1) or Me<sub>2</sub>CO. The *picrate* has *m* 185-185°. [Jampolsky et al. *J Am Chem Soc* 74 5222 1952, Chow & Fouss *J Am Chem Soc* 80 1095 1958, *Beilstein* 23 III/IV 1400.]

**2,5-Distyrylpyrazine** [14990-02-4] *M* 284.3, *m* 219°. Recrystallise it from xylene; chromatograph it on basic silica gel (60-80 mesh) using CH<sub>2</sub>Cl<sub>2</sub> as eluent, then sublime it in a vacuum on to a cold surface at 10<sup>-3</sup> torr [Ebied *J Chem Soc, Faraday Trans 1* 78 3213 1982]. All operations should be carried out in the dark.

**1,3-Dithiane** [505-23-7] *M* 120.2, *m* 54°. Crystallise the 1,3-dithiane from 1.5 times its weight of MeOH at 0°, and sublime it at 40-50°/0.1mm. [Gröbel & Seebach *Synthesis* 357 1977, *Beilstein* 19/1 V 13.]

**2,2'-Dithiobis(benzothiazole)** [120-78-8] *M* 332.2, *m* 180°, 182.5-183.5°. Crystallise it from \*benzene. [*Beilstein* 27 H 109, 27 III/IV 1862.]

**4,4'-Dithiodimorpholine (S,S'-di-N,N'-dimorpholine, dimorpholine-N,N'-disulfide)** [103-34-4] *M* 236.2, *m* 124-125°. Crystallise it from hot aqueous dimethylformamide or EtOH. It is a fungicide. [Blake *J Am Chem Soc* 65 1267 1943.]

**1-Dodecylpyridinium chloride** [104-74-5] *M* 301.9, *m* 68-70°. Purify the chloride by repeated crystallisation from acetone (charcoal); then recrystallise it twice from EtOH [Chu & Thomas *J Am Chem Soc* 108 6270 1986]. It is *hygroscopic* and should be stored with a desiccant. [*Beilstein* 20 III/IV 2314.]

**Ellipticine (5,11-dimethylpyrido[4,3-*b*]carbazole)** [519-23-3] *M* 246.3, *m* 311-315°(dec), 312-314°(dec),  $pK_{25} 5.78$  (80% aqueous methoxyethanol). This DNA intercalator is purified by recrystallisation from CHCl<sub>3</sub> or MeOH and is dried *in vacuo*. The UV has  $\lambda_{\text{max}}$  values in aqueous EtOH/HCl at 241, 249, 307, 335 and 426nm. [Marini-Bettolo & Schmutz *Helv Chim Acta* 42 2146 1959.] The *methiodide* has *m* 360°(dec), with UV  $\lambda_{\text{max}}$  (EtOH/KOH) at 223, 242, 251, 311, 362 and 432nm. [Goodwin et al. *J Am Chem Soc* 81 1903 1959, *Beilstein* 23/9 V 417.]

**Elymoclavine (8,9-didehydro-6-methylergoline-8-methanol)** [548-43-6] *M* 254.3, *m* 249-253°(dec), 250-252°(dec),  $[\alpha]_{\text{D}}^{20} -109^{\circ}$  (c 0.4, EtOH). It crystallises from MeOH, CHCl<sub>3</sub>, Et<sub>2</sub>O, Me<sub>2</sub>CO or \*C<sub>6</sub>H<sub>6</sub>. [*Beilstein* 23 III/IV 2716.]

**Emetine hydrochloride hydrate** [316-42-7] *M* 553.6 + aq, *m* 235-240°, 235-250°, 240-250°, 248-250° (depending on H<sub>2</sub>O content),  $[\alpha]_{\text{D}}^{20} -49.2^{\circ}$  (free base, c 4, CHCl<sub>3</sub>), +18° (c 6, H<sub>2</sub>O, dry salt),  $pK_1 5.77$ ,  $pK_2 6.64$ . It crystallises from MeOH/Et<sub>2</sub>O, MeOH or Et<sub>2</sub>O/EtOAc. The *free base* has *m* 104-105°, and the (-)-*phen,thiourea* derivative has *m* 220-221° (from EtOAc petroleum ether,  $[\alpha]_{\text{D}}^{25} -29.3^{\circ}$  (CHCl<sub>3</sub>)). Its IR has  $\nu_{\text{max}}$  at 3413 (OH) and 2611 (NH<sup>+</sup>) cm<sup>-1</sup>; and UV  $\lambda_{\text{max}}$  at 230nm ( $\epsilon$  16 200) and 282nm ( $\epsilon$  6 890) [Brossi et al. *Helv Chim Acta* 42 1515 1959, Barash et al. *J Chem Soc* 3530 1959]. [*Beilstein* 23 III/IV 3419.]

**Ergocornine** [564-36-3] *M* 561.7, *m* 182-184°,  $[\alpha]_{\text{D}}^{20} -176^{\circ}$  (c 0.5, CHCl<sub>3</sub>). It crystallises with solvent of crystallisation from MeOH. [Stadler et al. *Helv Chim Acta* 52 1549 1969, *Beilstein* 25 III IV 963, 27 II 860.]

**Ergocristine** [511-08-0] *M* 573.7, *m* 165-170°,  $[\alpha]_{\text{D}}^{20} -183^{\circ}$  (c 0.5, CHCl<sub>3</sub>). It crystallises with 2 molecules of solvent of crystallisation from \*benzene. [Stadler et al. *Helv Chim Acta* 52 1549 1969, *Beilstein* 25 III IV 966, 27 II 860.]

**Ergocryptine** [511-09-1] *M* 575.7, *m* 212-214°,  $[\alpha]_{\text{D}}^{20} -180^{\circ}$  (c 0.5, CHCl<sub>3</sub>). It crystallises with solvent of crystallisation, from acetone, \*benzene or methanol. [Stadler et al. *Helv Chim Acta* 52 1549 1969, *Beilstein* 25 III/IV 964, 27 II 860.]

**Ergotamine** [113-15-5] M 581.6, m 212-214°(dec),  $[\alpha]_D^{20}$  -160° (c 0.5, CHCl<sub>3</sub>), pK<sup>25</sup> 6.40. Crystallise it from \*benzene, then dry it by prolonged heating in high vacuum. It is very *hygroscopic*. [Beilstein 25 III/IV 964.]

**Ergotamine tartrate** [379-79-3] M 657.1, m 203°(dec). It crystallises from MeOH. [Beilstein 25 III/IV 964.]

**Ergotaminine** [639-81-6] M 581.7, m 241-243°,  $[\alpha]_D^{20}$  +369° (c 0.5, CHCl<sub>3</sub>). It forms rhombic plates from MeOH which retain solvent unlike its isomer ergotamine (previous entry). It is less soluble than ergotamine, and its solubility is 0.1% in boiling ethanol and 0.07% in methanol. [Stoll *Helv Chim Acta* 28 1299 and 1306 1945, Beilstein 25 II 860, 862, 25 III/IV 966.]

**D-Erythronic acid (3R-3,4-dihydroxyfuran-2-one)** [15667-21-7] M 118.1, m 98-100°, 103-104°, 104-105°, 105°,  $[\alpha]_D^{20}$  -73.2° (c 0.5, H<sub>2</sub>O),  $[\alpha]_{546}^{20}$  -87.6° (c 4, H<sub>2</sub>O). Recrystallise it from EtOAc (20 parts) or isoPrOH (3 parts). [Baker & MacDonald *J Am Chem Soc* 82 230 1960, Glattfeld & Forbrich *J Am Chem Soc* 56 1209 1934, Weidenhagen & Wegner *Chem Ber* 72 2010 1939, Musich & Rapoport *J Am Chem Soc* 100 4865 1978, Beilstein 18/2 V 457.]

**Esculetin (cichorigenin, 6,7-dihydroycoumarin)** [305-01-1] M 178.2, m 272-275° (dec), 274° (dec), pK<sup>25</sup> 8.60 (70% aqueous EtOH), pK<sub>Est(1)</sub> ~8.7, pK<sub>Est(2)</sub> ~12.4. It forms prisms from AcOH, aqueous EtOH or aqueous MeOH, and provides leaflets on sublimation in a vacuum. [Kagan *J Am Chem Soc* 88 2617 1966, Marby et al. *Phytochemistry* 4 492 1965.] Esculin (the 6-glucoside) has m 215°(dec),  $[\alpha]_D^{20}$  -41° (c 5, pyridine). [Beilstein 18 III/IV 1322, 18/3 V 202.]

**Eserine (Physostigmine, Physostol, [(3aS-cis)-1,2,3,3a,8,8a-hexahydro-1,3a,8-trimethyl-pyrrolo[2,3-b]indol-5-ol methylcarbamate ester]** [57-47-6] M 275.4, m 102-104°, 105-106°,  $[\alpha]_D^{17}$  -67° (c 1.3, CHCl<sub>3</sub>),  $[\alpha]_D^{25}$  -120° (\*C<sub>6</sub>H<sub>6</sub>), pK<sub>1</sub><sup>25</sup> 1.96, pK<sub>2</sub><sup>25</sup> 8.08. Eserine crystallises from Et<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub> and forms an unstable low melting form m 86-87° [Harley-Mason & Jackson *J Chem Soc* 3651 1954, Wijnberg & Speckamp *Tetrahedron* 34 2399 1978]. [Beilstein 23/11 V 401.]

**2-Ethoxy-1-ethoxycarbonyl-1,2-dihydroquinoline (EEDQ)** [16357-59-8] M 247.3, m 63.5-65°, 66-67°. Dissolve EEDQ (~180g) in CHCl<sub>3</sub>, evaporate to dryness in a vacuum. Add dry Et<sub>2</sub>O (20ml) and a white solid separates on standing. Set aside for a few hours, collect the solid, wash it thoroughly with cold Et<sub>2</sub>O and dry it in a vacuum (~140g, m 63.5-65°). A further crop of solid (~25g) is obtained from the filtrate on standing overnight. [Fieser & Fieser *Reagents for Organic Synthesis* 2 191 1969, Belleau et al. *J Am Chem Soc* 90 823 1968 and 90 1651 1968, Beilstein 21/3 V 28.]

**Ethoxyquin (1,2-dihydro-6-ethoxy-2,2,4-trimethylquinoline)** [91-53-2] M 217.3, b 169°/12-13mm, d<sub>4</sub><sup>20</sup> 1.000, pK<sub>Est</sub> ~5.8. Purify Ethoxyquin by fractional distillation *in vacuo* whereby the distillate solidifies to a glass. [Knoevenagel *Chem Ber* 54 1723, 1730 1921.] The *methiodide* has m 179° (from EtOH), and the *1-phenylcarbamoyl* derivative has m 146-147° (from EtOH). [Beaver et al. *J Am Chem Soc* 79 1236 1957, Beilstein 21 III/IV 95.]

**2-Ethyl-1,2-benzisoxazolium tetrafluoroborate** [4611-62-5] M 235.0, m 107-109°, 109.5-110.2°. Recrystallise it from MeCN/EtOAc to give magnificent crystals. It is not *hygroscopic* but on long exposure to moisture it etches glass. It is light-sensitive and should be stored in brown glass bottles and free from moisture. The UV (H<sub>2</sub>O), has  $\lambda_{\max}$  at 258nm ( $\epsilon$  13 100) and  $\lambda_{\max}$  297nm ( $\epsilon$  2 900); the IR (CH<sub>2</sub>Cl<sub>2</sub>) has  $\nu_{\max}$  at 1613 (C=N) and 1111-1000 (BF<sub>4</sub><sup>-</sup>) [UV, IR, NMR: Kemp & Woodward *Tetrahedron* 21 3019 1965].

**N-Ethylcarbazole** [86-28-2] M 195.3, m 69-70°, 69-71°, b 199-200°/15mm. Recrystallise it from MeOH, EtOH, EtOH/water or isopropanol and dry it below 55°. [Beilstein 20 H 436, 20 II 282, 20 III/IV 3829.]

**Ethyl 1,3-dithiane-2-carboxylate** [20462-00-4] M 192.3, b 75-77°/0.2mm, 96°/0.4mm, d<sub>4</sub><sup>20</sup> 1.220, n<sub>D</sub><sup>25</sup> 1.5379. Dissolve the ester in CHCl<sub>3</sub>, wash with aqueous K<sub>2</sub>CO<sub>3</sub>, twice with H<sub>2</sub>O, dry over MgSO<sub>4</sub>, filter,

evaporate and distil the residue. [Eliel & Hartman *J Org Chem* **37** 505 1972, Seebach *Synthesis* **1** 17 1969, *Beilstein* **19/7** V 227.]

**Ethyl 1,3-dithiolane-2-carboxylate** [20461-99-8] **M 178.3**, **b 85°/0.1mm**,  $d_4^{20}$  **1.250**,  $n_D^{20}$  **1.538**. Dissolve the ester in  $\text{CHCl}_3$ , wash it with aqueous  $\text{K}_2\text{CO}_3$ , twice with  $\text{H}_2\text{O}$ , dry it over  $\text{MgSO}_4$ , filter, evaporate and distil the residue *in vacuo*. [Hermann et al *Tetrahedron Lett* 2599 1973, Corey & Erickson *J Org Chem* **36** 3553 1971]. [*Beilstein* **19/7** V 225.]

**Ethylene oxide (oxirane)** [75-21-8] **M 44.0**, **b 13.5°/746mm**,  $d_4^{10}$  **0.882**,  $n_D^7$  **1.3597**. Dry oxirane with  $\text{CaSO}_4$ , then distil it from crushed  $\text{NaOH}$ . It has also been purified by passage, as a gas, through towers containing solid  $\text{NaOH}$ . [*Beilstein* **17/1** V 3.]

**Ethylene thiourea (2-imidazolidinethione)** [96-45-7] **M 102.2**, **m 203-204°**. Crystallise it from  $\text{EtOH}$  or amyl alcohol. [*Beilstein* **24** III/IV 22.]

**Ethylene urea (2-imidazolidone)** [120-93-4] **M 86.1**, **m 131°**. Crystallise it from  $\text{MeOH}$  (charcoal). [*Beilstein* **24** III/IV 6.]

**(±)-2-Ethylethylenimine (2-ethylaziridine)** [2549-67-9] **M 71.1**, **b 88.5-89°**,  $\text{pK}_{\text{a}}^{25}$  **8.31** ( $K_{\text{b}}$   $5.70 \times 10^{-7}$ ). Freshly distil the aziridine from sodium before use. The *picrate* has **m 103-104°**. **TOXIC**. [O'Rourke et al. *J Am Chem Soc* **78** 2159 1956, *Beilstein* **20** III/IV 280.]

**Ethyl hydrocupreine hydrochloride (Optochin)** [3413-58-9] **M 376.9**, **m 249-251°**,  $\text{pK}_{\text{a}}^{25}$  **5.50**,  $\text{pK}_{\text{a}}^{25}$  **9.95**. Recrystallise it from  $\text{H}_2\text{O}$  [UV: Heidt & Forbes *J Am Chem Soc* **55** 2701 1933]. [*Beilstein* **24** H 385, **24** III/IV 1446.]

**2-Ethyl-isothionicotinamide (ethionamide, 3-ethyl 4-pyridinecarbothioamide)** [536-33-4] **M 166.2**, **m 163-164°**, **164-166°(dec)**. It crystallises from  $\text{EtOH}$  as lemon yellow needles. The *hydrochloride* crystallises from  $\text{EtOH}$  (+ few drops of  $\text{HCl}$ ) as orange yellow needles with **m 212-214°**. [Kutscherowa et al. *J Gen Chem USSR* (English transl) **29** 915 1959, *Beilstein* **22** III/IV 737.] It causes peripheral and ocular **neuropathy** and is **carcinogenic** and **teratogenic**.

**(±)-3-Ethyl-5-phenylhydantoin (Ethotoin)** [86-35-1] **M 204.2**, **m 94°**. Crystallise it from water. It is an anticonvulsant and is used in epilepsy. [Dudley et al. *J Heterocycl Chem* **10** 173 1973, Pinner *Chem Ber* **21** 2320 1888, *Beilstein* **25** III/IV 963, **27** II 860.]

**N-Ethyl-5-phenylisoxazolinium-3'-sulfonate (Woodward's reagent K)** [4156-16-5] **M 253.3**, **m 220°(dec)**. Crystallise the reagent from diethyl ether or ethyl acetate/petroleum ether. [Lamas et al. *J Am Chem Soc* **108** 5543 1986.] It is best purified by dissolving in excess of aqueous  $\text{N HCl}$  and precipitating with  $\text{Me}_2\text{CO}$  to give a white fluffy solid. [Woodward et al. *Tetrahedron* **22** Suppl **8** 321 1966, Fieser & Fieser *Reagents for Organic Synthesis* **1** 385, **2** 198.]

**(±)-3-Ethyl-3-phenyl-2,6-piperidinedione (Glutethimide)** [77-21-4] **M 217.3**, **m 84°**. Crystallise glutethimide from diethyl ether or ethyl acetate/petroleum ether. It has **m 91-92°** (from aqueous  $\text{EtOH}$ ), **87-87.5°** (from  $\text{Et}_2\text{O}$ /petroleum ether), **84-87°** (from isopropanol), and **83-84°** (from  $\text{Et}_2\text{O}$ ). [Penprase & Biles *J Am Pharm Assoc* **47** 523 1958, Hofmann et al. *Helv Chim Acta* **40** 387, 393 1957, *Beilstein* **21** III/IV 5493.] The *R(+)-enantiomer* crystallises from  $\text{EtOAc}$ /petroleum ether with **m 103-104°**, and  $[\alpha]_D^{20} +184^\circ$  (c 1,  $\text{EtOH}$ ). [Branchini *Ricerche Scientifiche* **29** 2435 1959, Finch et al. *Experientia* **31** 1002 1975.]

**2-Ethylpyridine** [100-71-0] **M 107.2**, **b 148.6°**,  $d_4^{25}$  **0.942**,  $\text{pK}_{\text{a}}^{25}$  **5.89**. Dry 2-ethylpyridine with  $\text{BaO}$ , and fractionally distil it. Purify it further by conversion to the *picrate*, recrystallisation of the *picrate* and regeneration of the free base followed by distillation. [*Beilstein* **20/6** V 3.]

**4-Ethylpyridine** [536-75-4] **M 107.2, b 168.2-168.3°, d<sub>4</sub><sup>25</sup> 0.942, pK<sup>25</sup> 6.02.** Dry 4-ethylpyridine with BaO, and fractionally distil it. Also purified by converting to the *picrate*, recrystallising and the free base is regenerated and distilled. [Beilstein 20/6 V 10.]

**4-Ethylpyridine-1-oxide** [14906-55-9] **M 123.1, m 109-110°, pK<sub>Est</sub> ~1.1.** Crystallise the oxide from acetone/ether. [Beilstein 20/6 V 10.]

**Etioporphyrin I** See aetioporphyrin I above.

**Flavone (2-phenyl-4H-1-benzopyran-4-one)** [525-82-6] **M 222.3, m 100°.** Crystallise flavone from petroleum ether. [Beilstein 17/10 V 552.]

**Fluorescein [9-(*o*-carboxyphenyl-6-hydroxy-3H-xanthene-3-one)]** [2321-07-5] **M 320.0, ε<sub>495nm</sub> 7.84 x 10<sup>4</sup> (in 10<sup>-3</sup>M NaOH), pK<sub>1</sub> 2.2, pK<sub>2</sub> 4.4, pK<sub>3</sub> 6.7.** Dissolve it in dilute aqueous NaOH, filter and precipitate it by adding dilute (1:1) HCl. The process is repeated twice more, and the fluorescein is dried at 100°. *Alternatively*, it has been crystallised from acetone by allowing the solution to evaporate at 37° in an open beaker. It has also been recrystallised from EtOH and dried in a vacuum oven. [Beilstein 19 I 721, 19 II 248, 19 III/IV 2904, 19/8 V 456.]

**Fluoresceinamine (mixture of 5- and 6-aminofluorescein)** [27599-63-9] **M 347.3, m 215-220°(dec, 5-amino) and m 314-316°(dec, 6-amino).** Dissolve it in EtOH, treat with charcoal, filter, evaporate and dry the residue in a vacuum at 100° overnight. Also recrystallise it from 6% HCl, then dissolve it in 0.5% aqueous NaOH and precipitate it by acidifying with acetic acid. The separate amines are made from the respective nitro compounds, which are best separated *via* their acetate salts. They have similar R<sub>F</sub> of 0.26 on Silica Gel Merck F<sub>254</sub> in 5 ml MeOH + 150 ml Et<sub>2</sub>O saturated with H<sub>2</sub>O. The IR (Me<sub>2</sub>SO) has a band at ν<sub>max</sub> 1690 cm<sup>-1</sup> (CO<sub>2</sub><sup>-</sup>) and sometimes a weak band at ν<sub>max</sub> 1750 cm<sup>-1</sup> due to the lactone. The UV (EtOH) of the 6-isomer has λ<sub>max</sub> at 222nm (ε 60,000) and the 5-isomer at λ<sub>max</sub> 222nm (ε 60,000) and 285nm (ε 20,600). [IR: McKinney & Churchill *J Chem Soc (C)* 654 1970, McKinney et al. *J Org Chem* 27 3986 1962, UV: Verbiscar *J Org Chem* 29 490 1964, Beilstein 19 III/IV 4337, 19/8 V 713.]

**Fluorescein isothiocyanate isomer I (5-isothiocyanato isomer)** [3326-32-7; 27072-45-3 *mixture of 5- and 6-isomers*] **M 389.4, m >160°(slow dec).** It is made from the pure 5-amino isomer. Purify it by dissolving it in boiling Me<sub>2</sub>CO, filtering and adding petroleum ether (b 60-70°) until it becomes turbid. If an oil separates, then decant it and add more petroleum ether to the supernatant and cool. Orange-yellow crystals separate, collect and dry them *in vacuo*. It should give one spot on TLC (silica gel) in EtOAc/pyridine/AcOH (50:1:1) and in Me<sub>2</sub>NCHO/CHCl<sub>3</sub>/28% NH<sub>4</sub>OH (10:5:4). Its IR (Me<sub>2</sub>SO) has ν<sub>max</sub> at 2110 (NCS) and 1760 (C=O) cm<sup>-1</sup>. The <sup>1</sup>HNMR spectra in Me<sub>2</sub>CO-d<sub>6</sub> of the 5- and 6-isomers are distinctly different for the protons in the \*benzene ring; the UV in phosphate buffer pH 8.0 shows a λ<sub>max</sub> at ~490nm. [Sinsheimer et al. *Anal Biochem* 57 227 1974, McKinney et al. *Anal Biochem* 7 74 1964, Beilstein 19 III/IV 4337.]

**3-Fluoro-4-iodopyridine** [22282-75-3] **M 223.0, m 80-81°, 85-89°, pK<sub>Est</sub> ~1.7.** Crystallise it from petroleum ether and/or sublime it *in vacuo* (m 87°). The *picrate* [22282-76-4] has m 140° (from EtOH). [Gribble & Saulnier *Tetrahedron Lett* 4137 1980.]

**4-Fluoro-7-nitrobenzofurazan (4-fluoro-7-nitrobenzo-2-oxa-1,3-diazole)** [29270-56-2] **M 183.1, m 52.5-53.5°, 53-56°, 53.5-54.5°.** Purify it by repeated recrystallisation from petroleum ether (b 40-60°). On treatment with MeONa in MeOH it gives *4-methoxy-7-nitrobenzo-2-oxa-1,3-diazole* m 115-116°. [Nunno et al. *J Chem Soc (C)* 1433 1970.] It is a very good fluorophore for amino acids [Imai & Watanabe *Analyt Chim Acta* 130 377 1981], as it reacts with primary and secondary amines to form fluorescent adducts with λ<sub>ex</sub> 470nm and λ<sub>em</sub> 530nm. It gives a *glycine* derivative with m 185-187° [Miyano et al. *Anal Chim Acta* 170 81 1985].

**5-Fluorouracil (5-fluoropyrimidinedi-2,4-[1H,3H]-one, 5-FU)** [51-21-8] **M 130.1, m 282-283° (dec),**

**282-286°(dec), pK<sub>1</sub><sup>25</sup> 8.04, pK<sub>2</sub><sup>25</sup> 13.0.** Recrystallise it from H<sub>2</sub>O or MeOH/Et<sub>2</sub>O and sublime it at 190-200°/0.1mm or 210-230°/0.5mm. UV: λ<sub>max</sub> at 265-266nm (ε 7070). [Barton et al. *J Org Chem* **37** 329 1972, Duschinsky & Plevin *J Am Chem Soc* **79** 4559 1957, *Beilstein* **24** III/IV 1229.]

**Fluram (Fluorescamine, 4-phenyl-spiro[furan-2(3H)-1-phthalan]-3,3'-dione)** [38183-12-9] **M 278.3, m 153-155°, 154-155°.** Fluram is a non-fluorescent reagent that reacts with primary amines to form highly fluorescent compounds. Purify it by dissolving (~1g) in Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> (1:1, 180 ml), washing with 1% aqueous NaHCO<sub>3</sub> (50ml), drying (Na<sub>2</sub>SO<sub>4</sub>), and evaporating in a vacuum. Dissolve the residue in warm CH<sub>2</sub>Cl<sub>2</sub> (5ml), dilute with Et<sub>2</sub>O (12ml) and refrigerate. Collect the solid and dry it in a vacuum. Its IR (CHCl<sub>3</sub>) has ν<sub>max</sub> at 1810, 1745, 1722, 1625 and 1600 cm<sup>-1</sup>, and <sup>1</sup>HNMR (CDCl<sub>3</sub>) with δ at 8.71 (s, -OHC=). [Weigele et al. *J Am Chem Soc* **94** 5927 1972, Weigele et al. *J Org Chem* **41** 388 1976, *Lai Methods Enzymol* **47** 236 1977.]

**Forskolin (Colforsin, Coleonol, 5-[acetyloxy]-3-ethenyldodecahydro-6,10,10b-trihydroxy-3,4a,7,7, 10a-penta-methyl-[3R-{3α-4aβ, 5β, 6β, 6α, 10α, 10 aβ, 10bα}-1H-naphtho[2,1-b]pyran-1-one)** [66575-29-9] **M 410.5, m 230-232°, 228-233°, [α]<sub>D</sub><sup>25</sup> -96.2° (c 1.7, CHCl<sub>3</sub>).** Recrystallise it from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether, EtOAc/petroleum ether. It is an antihypertensive, a positive inotropic, a platelet aggregation inhibitor, and it has adenylate cyclase activating properties [*Chem Abstr* **89** 244150 1978, de Souza et al. *Med Res Rev* **3** 201 1983, X-ray: Tandon et al. *Indian J Chem* **15B** 880 1977]. [*Beilstein* **18/5** V 55.]

**Fumagillin {2.4.6.8-decatetraene-1,10-dioic acid mono[4-(1,2-epoxy-1,5-dimethyl-4-hexenyl)-5-methoxy-1-oxaspiro[2.5]oct-6-yl] ester}** [23110-15-8] **M 458.5, m 194-195°, [α]<sub>D</sub><sup>20</sup> -26.2° (in 95% EtOH), pK<sub>Est</sub> ~4.5.** Forty grams of a commercial sample containing 42% fumagillin, 45% sucrose, 10% antifoam agent and 3% of other impurities are digested with 150ml of CHCl<sub>3</sub>. The insoluble sucrose is filtered off and washed with CHCl<sub>3</sub>. The combined CHCl<sub>3</sub> extracts are evaporated almost to dryness at room temperature under reduced pressure. The residue is triturated with 20ml of MeOH, and the fumagillin is filtered off by suction. It is crystallised twice from 500ml of hot MeOH by standing overnight in a refrigerator (yellow needles). (The long-chain fatty ester used as antifoam agent is still present, but is then removed by repeated digestion, on a steam bath, with 100ml of diethyl ether.) For further purification, the fumagillin (10g) is dissolved in 150ml of 0.2M ammonia, and the insoluble residue is filtered off. The ammonia solution (cooled in running cold water) is then brought to pH 4 by careful addition of M HCl with constant shaking in the presence of 150ml of CHCl<sub>3</sub>. (Fumagillin is acid-labile and must be removed rapidly from the aqueous acid solution.) The CHCl<sub>3</sub> extract is washed several times with distilled water, dried (Na<sub>2</sub>SO<sub>4</sub>) and evaporated under reduced pressure. The solid residue is washed with 20ml of MeOH. The fumagillin is filtered off by suction, then crystallised from 200ml of hot MeOH. [Tarbell et al. *J Am Chem Soc* **77** 5610 1955.]

*Alternatively*, 10g of fumagillin in 100ml CHCl<sub>3</sub> is passed through a silica gel (5g) column to remove tarry material, and the CHCl<sub>3</sub> is evaporated to leave an oil which gives fumagillin on crystallisation from amyl acetate. It recrystallises from MeOH (charcoal) or Me<sub>2</sub>CO/MeOH. The fumagillin is stored in dark bottles in the absence of oxygen and at low temperatures. [Schenk et al. *J Am Chem Soc* **77** 5606 1955, *Beilstein* **19** III/IV 1012.]

**Furan** [110-00-9] **M 68.1, b 31.3°, d<sub>4</sub><sup>20</sup> 1.42, n<sub>D</sub><sup>20</sup> 1.4214.** Shake it with aqueous 5% KOH, dry it with CaSO<sub>4</sub> or Na<sub>2</sub>SO<sub>4</sub>, then distil it under nitrogen, from KOH or sodium, immediately before use. A trace of hydroquinone could be added as an inhibitor of oxidation. [*Beilstein* **17** H 27, **17** I 16, **17** II 34, **17** I V 291.]

**Furan-2-carboxylic (2-furoic) acid** [88-14-2] **M 112.1, m 133-134°, b 141-144°/20mm, 230-232°/760mm, pK<sub>1</sub><sup>25</sup> -7.3 (O-protonation), pK<sub>2</sub><sup>25</sup> 3.32.** Crystallise the acid from hot water (charcoal), dry it at 120° for 2 hours, then recrystallise it from CHCl<sub>3</sub>, and again dry it at 120° for 2 hours. For use as a standard in volumetric analysis, good quality commercial acid should be crystallised from CHCl<sub>3</sub> and dried as above or sublimed at 130-140°/50-60mm or less. [*Beilstein* **18** I 438, **18** II 265, **18** III/IV 3914, **18/6** V 102.]

**Furan-3-carboxylic (3-furoic) acid** [488-93-7] **M 112.1, m 122-123°, 123°.5, pK<sub>25</sub> 4.03.** Crystallise the acid from water or aqueous EtOH, and sublime it in a vacuum. [*Beilstein* **18** I 439, **18** III/IV 4052, **18/6** V 196.]

**Furan-3,4-dicarboxylic acid** [3387-26-6] **M 156.1, m 217-218°, 221.5-222.5°, pK<sub>1</sub><sup>25</sup> 1.44, pK<sub>2</sub><sup>25</sup> 7.84.**

Crystallise it from water or Et<sub>2</sub>O/petroleum ether, and sublime it in a vacuum. [*Beilstein* 18 III/IV 4497.]

**Furan-2,5-dione (maleic anhydride)** [108-31-6] **M 98.1, m 54°, b 94-96°/20mm, 199°/760mm.** Crystallise it from \*benzene, CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub> or CCl<sub>4</sub>. Sublime it under reduced pressure. [Skell et al. *J Am Chem Soc* 108 6300 1986, *Beilstein* 17 III/IV 5897, 17/11 V 55.]

**3-(2-Furanyl)acrylic acid** [539-47-9] **M 138.1, (cis-isomer) m 106-108°, pK<sup>25</sup> 3.5; (trans-isomer) m 141°, 143-144°, b 285°, pK<sup>25</sup> 4.5 (H<sub>2</sub>O), 5.76 (50% aqueous EtOH), 6.65 (ethoxyethanol/H<sub>2</sub>O—80:20).** Crystallise the *cis*-isomer from \*C<sub>6</sub>H<sub>6</sub> and the *trans*-isomer from H<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> or petroleum ether (b 80-100°)(charcoal). [*Beilstein* 18 H 301, 18 III/IV 4143, 18/6 V 306.]

**Furfural (2-furfuraldehyde)** [98-01-1] **M 96.1, b 54-56°/11mm, 59-60°/15mm, 67.8°/20mm, 90°/65mm, 161°/760mm, d<sub>4</sub><sup>20</sup> 1.159, n<sub>D</sub><sup>20</sup> 1.52608, pK<sup>25</sup> -6.5 (O-protonation).** Furfural is unstable to air, light and acids. Impurities include formic acid, β-formylacrylic acid and furan-2-carboxylic acid. Distil it in an oil bath from 7% (w/w) Na<sub>2</sub>CO<sub>3</sub> (added to neutralise acids, especially pyromucic acid). Redistil it from 2% (w/w) Na<sub>2</sub>CO<sub>3</sub>, and then, finally fractionally distil it under vacuum. It is stored in the dark. [Evans & Aylesworth *Ind Eng Chem (Anal ed)* 18 24 1926.]

Impurities resulting from storage can be removed by passage through chromatographic grade alumina. Furfural can be separated from impurities other than carbonyl compounds by the bisulfite addition compound. The aldehyde is steam volatile.

It has been purified by distillation (using a Claisen head) under reduced pressure. This is essential as is the use of an oil bath with temperatures of no higher than 130° which is highly recommended. When furfural is distilled at atmospheric pressure (in a stream of N<sub>2</sub>), or under reduced pressure with a free flame (caution: because the aldehyde is flammable), an almost colourless oil is obtained. After a few days and sometimes a few hours, the oil gradually darkens and finally becomes black. This change is accelerated by light and occurs more slowly when it is kept in a brown bottle. However, when the aldehyde is distilled under vacuum and the bath temperature kept below 130° during the distillation, the oil develops only a slight colour when exposed to direct sunlight during several days. The distillation of very impure material should NOT be attempted at atmospheric pressure; otherwise the product darkens very rapidly. After one distillation under vacuum, a distillation at atmospheric pressure can be carried out without too much decomposition and darkening. The liquid **irritates mucous membranes**. Store it in dark containers under N<sub>2</sub>, preferably in sealed ampoules. [Adams & Voorhees *Org Synth Coll Vol I* 280 1941, *Beilstein* 17/9 V 292.]

**Furfuryl alcohol (2-furyl-methanol)** [98-00-0] **M 98.1, b 68-69°/20mm, 170.0°/750mm, d<sub>4</sub><sup>20</sup> 1.132, n<sub>D</sub><sup>20</sup> 1.4873, n<sub>D</sub><sup>30</sup> 1.4801, pK<sup>25</sup> 2.61.** Distil it under reduced pressure to remove tarry material, shake with aqueous NaHCO<sub>3</sub>, dry it with Na<sub>2</sub>SO<sub>4</sub> and fractionally distil it under reduced pressure from Na<sub>2</sub>CO<sub>3</sub>. It can be further dried by shaking with Linde 5A molecular sieves. [*Beilstein* 17/3 V 338.]

**Furfurylamine (2-aminomethylfuran)** [617-89-0] **M 97.1, b 54-56°/17mm, 142.5-143°/735mm, d<sub>4</sub><sup>20</sup> 1.059, n<sub>D</sub><sup>20</sup> 1.489, pK<sup>30</sup> 8.89.** Distil it under nitrogen from KOH through a column packed with glass helices, preferably under vacuum. The *picrate* has **m** 184-184°(dec), the *hydrochloride* has **m** 147-149°, and the *oxalate* has **m** 145-147°. [*Beilstein* 18 H 584, 18 II 416, 18 III/IV 3068, 18/9 V 541.]

**6-Furfurylaminopurine (Kinetin)** [525-79-1] **M 215.2, m 266-267°, 269-271°, 270-272°, 272° (sealed capillary), pK<sub>1</sub> <1, pK<sub>2</sub> 3.8, pK<sub>3</sub> 10.** It forms platelets from EtOH and sublimes at 220°, but is best done at lower temperatures in a good vacuum. It has been extracted from neutral aqueous solutions with Et<sub>2</sub>O. [Miller et al *J Am Chem Soc* 78 1375 1956, Bullock et al. *J Am Chem Soc* 78 3693 1956, *Beilstein* 26 III/IV 3586.]

**Furil** [492-94-4] **M 190.2, m 165-166°.** Furil crystallises from MeOH or \*benzene (charcoal). [*Beilstein* 19 III/IV 2008.]

**(±)-Furoin [1,2-di-(2-furyl)-2-hydroxyethanone]** [552-86-3] **M 192.2, m 135-136°, 138-139°, 158-162°/9mm.** It crystallises from MeOH (charcoal) and distils in a vacuum. [Hartman & Dickey *J Am Chem Soc* 55 1228 1933.] The (-)-*enantiomer* crystallises from toluene or EtOH with **m** 131-131°, and [α]<sub>D</sub><sup>20</sup> -4.9°

(dioxane) [Neuberg et al. *Arch Biochem* **1** 393 1943, *Beilstein* **19** H 204, **19** I 710, **19** II 224, **19** III/IV 2543.]

**Fusaric acid (5-*n*-butylpyridine-2-carboxylic acid)** [536-69-6] **M 179.2, m 96-98°**, **98°**, **98-100°**, **101-103°**, **pK<sub>1</sub> 5.7, pK<sub>2</sub> 6.16 (80% aqueous methoxyethanol)**. Dissolve it in CHCl<sub>3</sub>, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and recrystallise the residue from 50 parts of petroleum ether (b 40-60°), CHCl<sub>3</sub>/petroleum ether or EtOAc, then sublime it *in vacuo*. The *amide* crystallises from MeOH with **m 128.2-129.0°**. The *copper salt* forms bluish violet crystals from H<sub>2</sub>O and has **m 258-259°**. [Hardegger & Nikles *Helv Chim Acta* **39** 505 1956, Schreiber & Adam *Chem Ber* **93** 1848 1960, NMR and MS: Tschesche & Führer *Chem Ber* **111** 3500 1978, *Beilstein* **22** III/IV 764, **22/2** V 384.]

**Fuschin (Magenta I)** See rosaniline hydrochloride in "Aromatic Compounds" [632-99-2] in this Chapter.

**Glycidol (oxirane-2-methanol)** [*RS*-(±)- 556-52-5; *R*-(+)- 57044-25-4; *S*-(-)- 60456-23-7] **M 74.1, (R,S) b 61-62°/15mm, d<sub>4</sub><sup>20</sup> 1.117, n<sub>D</sub><sup>20</sup> 1.433, [S(-)-isomer, § also available on polymer support, has b 49-50°/7mm, 66-67°/19mm, [α]<sub>D</sub><sup>20</sup> -15°(neat)], [R(+)-isomer has b 56-56.5°/11mm, d<sub>4</sub><sup>20</sup> 1.117, n<sub>D</sub><sup>20</sup> 1.429, [α]<sub>D</sub><sup>20</sup> +15°(neat)]**. Purify glycidol by fractional distillation. The *4-nitrobenzoates* have **m 56°(±)**; **m 60-62°**, [α]<sub>D</sub><sup>20</sup> -37.9° (c 3.38 CHCl<sub>3</sub>) for *R*-(-)-isomer [106268-95-5]; **m 60-62°**, [α]<sub>D</sub><sup>20</sup> +38° (c 1 CHCl<sub>3</sub>) for the *S*-(+)-isomer [115459-65-9] and are recrystallised from Et<sub>2</sub>O or Et<sub>2</sub>O/petroleum ether (b 40-60°) [*S*-isomer: Burgos et al. *J Org Chem* **52** 4973 1987, Sowden & Fischer *J Am Chem Soc* **64** 1291 1942.] [*Beilstein* **17** I 50, **17** III/IV 985, **17/3** V 9.]

**Gramine (3-dimethylaminoethylindole)** [87-52-5] **M 174.3, m 134°**, **pK<sub>25</sub> 16.00 (NH acidic)**, **basic pK<sub>25</sub> 9.2 (50% aqueous EtOH)**. Crystallise gramine from diethyl ether, ethanol or acetone. It sublimes at 59°/0.001mm. The *hydrochloride* crystallises from EtOH/Et<sub>2</sub>O with **m 190.5-191.0°(dec)**. [Culvenor et al. *Aust J Chem* **17** 1301 1964, *Beilstein* **22** III/IV 4302, **22/10** V 25.]

**(2*S*,6'*R*)(+)-Griseofulvin** [126-07-8] **M 352.8, m 220°**, [α]<sub>D</sub><sup>22</sup> +365° (c 1, acetone). Crystallise it from \*benzene or EtOH. Purify 2g of griseofulvin by chromatography on Alumina (40 x 1.5cm) and elute with \*C<sub>6</sub>H<sub>6</sub>/MeOH (199:1) and follow the UV blue fluorescent band. [MacMillan *J Chem Soc* 1823 1959, *Beilstein* **18** III/IV 3160, **18/5** V 150.]

**Guanosine (H<sub>2</sub>O)** [118-00-3] **M 283.2, m 237-237.5°(dec)**, **239°(dec)**, [α]<sub>546</sub><sup>20</sup> -86° (c 1, 0.1M NaOH), **pK<sub>1</sub><sup>25</sup> 1.9, pK<sub>2</sub><sup>25</sup> 9.24, pK<sub>3</sub><sup>25</sup> 12.33**. It crystallises from water as a *dihydrate*. Dry it at 110°. [*Beilstein* **26/18** V 81.]

**Guanylic acid (guanosine-5'-monophosphoric acid)** [85-32-5] **M 363.2, m 208°(dec)**, **pK<sub>2</sub><sup>25</sup> 2.4, pK<sub>3</sub><sup>25</sup> 6.66 (6.1), pK<sub>4</sub><sup>25</sup> 9.4**. Crystallise it from water and dry it at 110°. [*Beilstein* **26** III/IV 3910.]

**Harmaline (7-methoxy-1-methyl-4,9-dihydro-3*H*-β-carboline, 4,9-dihydro-7-methoxy-1-methyl-3*H*-pyrido[3,4-*b*]indole)** [304-21-2] **M 214.3, m 229-230°**, **229-231°**, **235-237° (after distillation at 120-140°/10<sup>-3</sup>)**, **pK<sub>1</sub> 4.2**. Recrystallise harmaline from MeOH and sublime it at high vacuum. It has UV in MeOH with λ<sub>max</sub> at 218, 260 and 376nm (log ε 4.27, 3.90 and 4.02, respectively); IR (Nujol) with ν<sub>max</sub> at 1620, 1600, 1570 and 1535cm<sup>-1</sup> and in CHCl<sub>3</sub> ν<sub>max</sub> at 1470 and 1629cm<sup>-1</sup>. [Spenser *Can J Chem* **37** 1851 1959, Marion et al. *J Am Chem Soc* **73** 305 1951, UV Prukner & Witkop *Justus Liebigs Ann Chem* **554** 127 1942.] The *hydrochloride dihydrate* has **m 234-236°(dec)**, the *picrate* has **m 228-229°** (sinters at 215°) from aqueous EtOH, and the *N-acetate* forms needles **m 204-205°**. [*Beilstein* **23** H 396, **23** I 119, **23** II 345, **23** III/IV 2666, **23/12** V 148.]

**Harmane (2-methyl-β-carboline, 1-methyl-9*H*-pyrido[3,4-*b*]indole, Aribine)** [486-84-0] **M 182.2, m 235-238°**, **237-238°**, **pK<sub>1</sub> 7.37 (basic, Pry N)**, **pK<sub>2</sub> 14.7 (acidic, NH)**. Crystallise it from heptane/cyclohexane. It is insoluble in H<sub>2</sub>O, but soluble in dilute HCl or H<sub>2</sub>SO<sub>4</sub>. Solutions show a blue fluorescence. Its UV (MeOH) has λ<sub>max</sub> nm (log ε) at 234 (4.57), 287 (4.21) and 347 (3.66). The *hydrochloride* forms needles from EtOH/dil HCl which sublimes at **m 120-130°**. It is an imidazoline binding site agonist. [Wolfbeis *Monatsh Chem* **113** 509 1982, *Beilstein* **23/12** 237.]

**Harmine (7-methoxy-1-methyl-9H-pyrido[3,4-b]indole)** [442-51-3] M 212.3, m 261°(dec), 265°, pK<sup>20</sup> 7.61. Crystallise harmine from MeOH and sublime it in a vacuum. Its UV (MeOH) has  $\lambda_{\text{max}}$  nm (log  $\epsilon$ ) at 241 (4.61), 301 (4.21) and 336 (3.69). It is a CNS stimulant. See *hydrochloride* below. [Beilstein 23 II 348, 23 III/IV 2702, 23/12 V 237.]

**Harmine hydrochloride (hydrate)** [343-27-1] M 248.7, m 262°(dec, hydrate), 280°(dec). The *hydrate* crystallises as fluorescent crystals from water. Freely soluble in hot water, but only 2.5% in cold water. The *anhydrous* salt has m 319°. [Beilstein 23/12 V 237.]

**Hesperetin (3',5,7-trihydroxy-4'-methoxyflavanone)** [520-33-2] M 302.3, (R,S) m 227-228°, pK<sub>Est</sub> ~8.5-10.5 (phenolic). Crystallise it from EtOAc or ethanol. The natural S(-) form crystallises from EtOH and has m 216-218° and  $[\alpha]_{\text{D}}^{20}$  -37.6° (c 2, EtOH). Note that C2 is chiral. [Beilstein 18 II 204, 18 III/IV 3215, 18/5 V 214.]

**Hesperidin (hesperetin 7-rhamnoside)** [520-26-3] M 610.6, m 258-262°, 261-263°(dec),  $[\alpha]_{546}^{20}$  -82° (c 2, pyridine). Dissolve hesperidine in dilute aqueous alkali and precipitate it by adjusting the pH to 6-7. [Beilstein 18 III/IV 3219, 18/5 V 218.]

**Hexahydro-1H-azepine (hexamethyleneimine, Azepane)** [111-49-9] M 99.2, b 70-72°/30mm, 135-138°/atm,  $d_4^{20}$  0.879,  $n_{\text{D}}^{20}$  1.466, pK<sup>25</sup> 11.10 (pK<sup>0</sup> 9.71, pK<sup>75</sup> 9.71). Purify azepane by dissolving in Et<sub>2</sub>O and adding ethanolic HCl until all the base separates as the white *hydrochloride*, filter, wash with Et<sub>2</sub>O and dry it (m 236°). The salt is dissolved in the minimum volume of H<sub>2</sub>O and basified to pH ~14 with 10N KOH. The solution is extracted with Et<sub>2</sub>O, the extract is dried over KOH, evaporated and distilled. The free base is a **FLAMMABLE** and **TOXIC** liquid, and best kept as the salt. The *nitrate* has m 120-123°, the *picrate* has m 145-147°, and the *tosylate* has m 76.5° (ligroin). [Müller & Sauerwald *Monatsh Chem* 48 727 1027, Hjelt & Agback *Acta Chem Scand* 18 194 1964, Beilstein 20 II 1406, 20 III/IV 1406, 20/4 V 3.]

**Hexamethylenetetramine (Urotropine, hexamine, HMTA)** [100-97-0] M 140.1, m 280° (subln), 290-292° (sealed tube, CARE),  $d_4^{20}$  1.331, pK<sup>25</sup> 4.85 (6.30). It is soluble in H<sub>2</sub>O (67%), CHCl<sub>3</sub> (10%), EtOH (8%) and Et<sub>2</sub>O (0.3%), and a 0.2M solution has a pH of 8.4. Dissolve it in hot absolute EtOH (reflux, Norit), filter using a heated funnel, cool at room temperature first, then in ice. Wash the crystals with cold Et<sub>2</sub>O, dry them in air or under a vacuum. A further crop can be obtained by adding Et<sub>2</sub>O to the filtrate. It sublimes above 260° without melting. The *picrate* has m 179°(dec). [pK<sup>20</sup> 4.85: Reilley & Schmid *Anal Chem* 30 947 1958, pK<sup>20</sup> 6.30: Pummerer & Hofmann *Chem Ber* 56 1255 1923.] [Beilstein 26 I 306, 26 II 200, 26 III/IV 1680.]

**R-(+)-2-Hexyloxirane [R-(+)-1,2-epoxyoctane, R-(+)-octene oxide]** [77495-66-0] M 128.2, b 51-54°/10mm, 60-62°/15mm, 157°/atm,  $d_4^{25}$  0.839,  $n_{\text{D}}^{20}$  1.418,  $[\alpha]_{\text{D}}^{20}$  +14° (neat), the S(-)- enantiomer has [50418-68-3]  $[\alpha]_{\text{D}}^{20}$  -14.5° (c 2.5, EtOH),  $n_{\text{D}}^{20}$  1.412. The enantiomeric oxiranes have been purified by passage through a silica gel column in pentane and eluted with pentane/Et<sub>2</sub>O. The eluate is evaporated and subjected to bulb-to-bulb distillation in a vacuum. The enantiomeric purity is checked by HPLC. [White & Emmons *Tetrahedron* 17 31 1962, Johnson & Rogers *J Org Chem* 38 1793 1973, Beilstein 17 H 17, 17 I 11, 17 III/IV 111, 17/1 V 138.]

**Histamine** [51-45-6] M 111.2, m 86° (sealed tube), b 167°/0.8mm, 209°/18mm, pK<sub>1</sub><sup>25</sup> 6.02, pK<sub>2</sub><sup>25</sup> 9.70. It crystallises from \*benzene or chloroform. [Beilstein 25 I 628, 25 II 302, 25 III/IV 2049.]

**Histamine dihydrochloride** [56-92-8] M 184.1, m 249-252° (244-245°). The dihydrochloride crystallises from aqueous EtOH. The *phosphate* (2H<sub>3</sub>PO<sub>4</sub>) [51-74-1] has m 132-133° (from H<sub>2</sub>O). [Beilstein 25 III/IV 2049.]

**Homopiperazine (1,4-diazepane)** [505-66-8] **M 100.2, m 38-40°, 43°, b 60°/10mm, 92°/50mm, 169°/atm, pK<sub>1</sub><sup>20</sup> 6.70, pK<sub>2</sub><sup>20</sup> 10.41.** Purify it by fractionation through a column of 10 theoretical plates with a reflux ratio of 3:1. It boils at 169°, and the cool distillate crystallises in plates **m 43°**. [Poppelsdorf and Myerly *J Org Chem* **26** 131 1961.] Its pKa values are 6.67 and 10.09 at 29.7°, and 6.28 and 9.86 at 40° [Pagano et al. *J Phys Chem* **65** 1062 1961]. The 1,4-bis(4-bromobenzoyl) derivative has **m 194-198°** (from EtOH); the hydrochloride has **m 270-290°** (from EtOH) and the picrate has **m 265°(dec)** [Lloyd et al. *J Chem Soc (C)* 780 1966]. [Beilstein **23** III/IV 388, **23/3** V 240.]

**1-Hydrazinophthalazine (hydralazine)** [86-54-4] **M 160.1, m 172-173°(dec), pK<sub>1</sub><sup>25</sup> 2.90, pK<sub>2</sub><sup>25</sup> 7.25 (NNH<sub>2</sub>).** It crystallises from MeOH. Its UV has  $\lambda_{\max}$  at 656nm at pH ~11. It complexes with Bi<sup>3+</sup>, Zn<sup>2+</sup>, Fe<sup>2+</sup> and Co<sup>2+</sup>. The hydrochloride (hydralazine hydrochloride) [304-20-1] **M 196.6**, also crystallises from MeOH and has **m 172-173°(dec)**. [Druey et al. *Helv Chim Acta* **34** 195 1951, *Beilstein* **25** III/IV 4552.] It is an antihypertensive.

**2-Hydrazinopyridine** [4930-98-7] **M 109.1, m 41-44°, 46-47°, 49-50°, b 105°/0.5mm, 128-135°/13mm.** Purify it by distillation under a vacuum and by recrystallisation from Et<sub>2</sub>O/hexane. [Kauffmann et al. *Justus Liebigs Ann Chem* **656** 103 1962, Potts & Burton *J Org Chem* **31** 251 1966.] The mono-hydrochloride has **m 183°(dec)** from aqueous HCl, and the di-hydrochloride has **m 214-215°**. [Beilstein **22** II 487, **22** III/IV 7025, **22/14** V 486.]

**4-Hydroxyacridine (4-acridinol, neo-oxine)** [18123-20-1] **M 195.2, m 116.5°, 122-123°, pK<sub>1</sub><sup>15</sup> 5.28, pK<sub>2</sub><sup>15</sup> 9.75.** Crystallise neo-oxine from EtOH or aqueous EtOH. It complexes with Zn<sup>2+</sup>, Cd<sup>2+</sup>, Ga<sup>2+</sup>, Pb<sup>2+</sup>, Cr<sup>2+</sup>, Mn<sup>2+</sup>, Fe<sup>2+</sup>, Co<sup>2+</sup> and Ni<sup>2+</sup>. The hydrochloride crystallises from EtOH with **m 242°**. [Beilstein **21** II 78, **21** III/IV 1562, **21/4** V 90.]

**1-Hydroxy-1,2-benziodoxol-3(1H)-one (IBX, 2-iodoxybenzoic acid)** [61717-82-6] **M 280.0, m 224-225°, 226-234°, 232-233°(dec), 233°(dec), pK <4.** IBX prepared by the Dess-Martin procedure (KBrO<sub>3</sub>/H<sub>2</sub>SO<sub>4</sub>) [Dess & Martin *J Am Chem Soc* **113** 7277 1991] has been reported as being explosive and comparable to TNT [Hartman & Meyer *Chem Ber* **26** 1727 1893, Plumb & Harper *Chem Eng News* (July 16) 3 1990] when apparently the iodine content was below 43.5% (theoretical value is 45.32%). In an attempt to avoid this, IBX has been washed with H<sub>2</sub>O and EtOH to render it non-explosive. Although deliberate attempts to detonate it on several occasions have been unsuccessful; operators should exercise CAUTION when working with IBX particularly with large scale preparations. This is among the substances for which it may be difficult to obtain permission to transport it, so detailed preparations are described here.

*Dess-Martin method:* To a vigorously stirred mixture of 2-iodobenzoic acid (85.2g, 340mmol, see [88-67-5]) and 0.73 M H<sub>2</sub>SO<sub>4</sub> (730ml) in a bath at 55° is added KBrO<sub>3</sub> (76.0g, 450mmol) over 1 hour. The mixture is stirred further for 3.6 hours at 65°, cooled in an ice bath, and the solid is filtered off, washed with H<sub>2</sub>O (1000ml) then EtOH (2 x 50ml) and dried *in vacuo* to give BTX (89.1g, 320mmol, 93%) of analytical purity. [Dess & Martin *J Am Chem Soc* **113** 7277 1991]

*Santagostino et al.'s method:* To oxone (181.0g, 290mmol, see [70693-62-8] 2KHSO<sub>5</sub>—KHSO<sub>4</sub>—K<sub>2</sub>SO<sub>4</sub> triple salt) in deionised H<sub>2</sub>O (650ml, 450mmol) is added rapidly 2-iodobenzoic acid (50.0g, 200mmol), and kept at 70-73° for 20 minutes then stirred mechanically at 70-73° for 3 hours. Initially a thick slurry coating along the walls of the container is formed which gradually becomes finely dispersed and the now easily stirred suspension readily sediments if stirring is interrupted. The mixture is then stirred slowly at 5° for 1.5 hours, the white crystalline solid is collected onto a medium porosity sintered-glass funnel, washed with H<sub>2</sub>O (6 x 100ml) and Me<sub>2</sub>CO (2 x 100ml) then Et<sub>2</sub>O, and left in dry air for 16 hours to give IBX (44.8-45.7g, ~80%). The filtrates are disposed of by treating with solid Na<sub>2</sub>SO<sub>3</sub> (70g, 55mmol) and neutralising with M aqueous NaOH. Analysis indicated that the purity was ≥95% (judging by the <sup>1</sup>H NMR integrals of the triplets at 7.84 and 7.47 ppm, and elemental analysis), and it contained 2-iodosobenzoic acid (~4%) and 2-iodobenzoic acid (~0.5%). IBX of analytical purity (≥99%) is similarly obtained on a small scale by adding 2-iodobenzoic acid (5.0g, 20mmol) to a solution of oxone in deionised H<sub>2</sub>O (37.2g, 61mmol, in 200ml), and the suspension is kept at 70° for 1 hour when it becomes clear. After 0.5 hours at 0-5° the white crystals that separate are collected, washed and dried as before to give analytically pure IBX (4.4g, 77%), **m 233° (dec)**. This procedure is more environmentally friendly than the Dess-Martin procedure. [Frigerio et al. *J Org Chem* **64** 4537 1999.] It has IR

(film) with a  $\nu_{\max}$  at  $1640\text{ cm}^{-1}$ ; the  $^1\text{H NMR}$  ( $\text{DMSO-}d_6$ , TMS) has  $\delta$  at 8.15 (d, 1H), 8.01 (d, 1H), 7.89 (t, 1H), and 7.84 (t, 1H); and the  $^{13}\text{C NMR}$  ( $\text{DMSO-}d_6$ , TMS) has  $\delta$  at 167.49, 146.59, 133.39, 132.97, 131.36, 130.10 and 124.99 [Frigerio & Santagostino *Tetrahedron Lett* **43** 8019 1994].

**SIBX** is an even more stabilised, non-explosive formulation that has all the oxidative characteristics of IBX. It is prepared as follows: 2-iodobenzoic acid (200g) and isophthalic acid (133g, see [121-91-5]) are added into a solution of oxone (625g) in  $\text{H}_2\text{O}$  (2000ml) and kept at  $70^\circ$  for 3 hours, then sodium benzoate (128g) in  $\text{H}_2\text{O}$  (500ml) is added at  $40^\circ$ . The precipitate that is formed on cooling to  $20^\circ$  is filtered off, washed with  $\text{H}_2\text{O}$  (700ml) and dried in a ventilated oven at  $60^\circ$  to give **SBTX** (420g, 90%) containing 49% w/w of active IBX. [Ozanne et al. *Org Lett* **5** 2903 2003.]

**SIBX** and/or **IBX** are water tolerant, inexpensive to prepare and easy to use. They oxidise primary and secondary alcohols smoothly to aldehydes and ketones, and vicinal diols are oxidised to mono and/or dicarbonyl compounds without C-C cleavage [Frigerio & Santagostino *Tetrahedron Lett* **43** 8019 1994], as well as allylic and benzylic alcohols to respective aldehydes [Ozanne et al. *Org Lett* **5** 2903 2003.], amino alcohols to amino carbonyls without amino protection, and sensitive heterocyclic substituents are not affected [Frigeio et al. *J Org Chem* **60** 7272 1995]. The structure and kinetics of the reactive intermediates in the oxidation have been studied in detail [Munari et al. *J Org Chem* **61** 9272 1961]. The reagents are generally sparingly soluble in organic solvents such as EtOAc, THF,  $\text{Me}_2\text{CO}$ , MeCN, toluene, DMSO and *N*-methylpyrrolidine, but the reactions usually proceed in these solvents, or mixtures of them, with ease and in good to excellent yields either at room temperature or at as high as reflux temperatures.

**2-Hydroxybenzothiazole (benzothiazol-2(3H)-one)** [934-34-9] **M 183.1, m 140-141 $^\circ$** . Crystallise it from aqueous EtOH or water. [Hoggarth *J Chem Soc* 3314 1949, Hunter *J Chem Soc* 135 1930, *Beilstein* **27** H 182, **27** I 270, **27** II 225, **27** III/IV 2693.]

**1-Hydroxybenzotriazole hydrate (HOBt)** [2592-95-2; 123333-53-9 ( $\text{H}_2\text{O}$ )] **M 135.1, m 157 $^\circ$ , 159-160 $^\circ$ , pK $^{20}$  7.88**. Crystallise HOBt from hot aqueous EtOH or water (charcoal). It is prepared from *o*-nitrophenylhydrazine or its hydrochloride dissolved in a small volume of  $\text{H}_2\text{O}$  by treating with 25% of aqueous KOH or aqueous ammonia whereby the mixture warms up immediately and discolours. After the mixture cools to room temperature it is acidified with hydrochloric acid and the *hydroxybenzotriazole* separates as colourless needles. It is soluble in hot  $\text{H}_2\text{O}$ , EtOH and AcOH, but much less soluble in  $\text{Et}_2\text{O}$ , petroleum ether,  $^*\text{C}_6\text{H}_6$  and  $\text{CHCl}_3$ . It is an acid and forms metal salts such the **Pb salt** which crystallises from hot  $\text{H}_2\text{O}$  in glistening leaflets with **m 270 $^\circ$** . In aqueous solution it exists mainly as the zwitterionic N-oxide tautomer with  $\text{H}^+$  on N-2. [Nietzki & Braunschweig *Chem Ber* **27** 3381 1894, Zincke & Schwartz *Justus Liebigs Ann Chem* **329** 329 1900, Boyle & Jones *J Chem Soc Perkin Trans II* 160 1973, Tomita & Ikawa *J Pharm Soc Jpn* **75** 449 1955, *Beilstein* **26** III/IV 95.] It is a useful reagent for peptide synthesis [Heusel et al. *Angew Chem Int Ed* **16** 642 1977].

§ A polystyrene supported version is available. For use in solid phase peptide synthesis, see Dryland & Sheppard *J Chem Soc Perkin Trans I* 125 1986.

**4-Hydroxycoumarin (4-hydroxy-1-benzopyran-2-one)** [1076-38-6] **M 162.1, m 206 $^\circ$ , 211-213 $^\circ$ , pK $_{\text{Est}}$  ~9.0**. Crystallise 4-hydroxycoumarin from water and dry it in a vacuum desiccator over Sicapent. [*Beilstein* **18/1** V 378.]

***N*-2-Hydroxyethylpiperazine-*N'*-2-ethanesulfonic acid (HEPES)** [7365-45-9] **M 238.3, pK $^{20}$  7.55**. Crystallise the acid from hot EtOH and water. It is a useful buffer. [*Beilstein* **23** V 376.]

**3-Hydroxyflavone (Flavanol)** [577-85-5] **M 238.2, m 169-170 $^\circ$ , 171-172 $^\circ$** . Recrystallise it from MeOH (**m** 169.5-170 $^\circ$ ), EtOH, aqueous EtOH (**m** 167 $^\circ$ ) or hexane. It has also been purified by repeated sublimation under high vacuum, and dried at high vacuum pumping for at least one hour [Bruker & Kelly *J Phys Chem* **91** 2856 1987]. [*Beilstein* **17** H 527, **17** I 268, **17** II 498, **17** III/IV 6428.]

**7-Hydroxy-4-methylcoumarin (4-methylumbelliferone( $\beta$ ) hydrate)** [90-33-5] **M 194.2, m 185-186 $^\circ$ , 188-188.5 $^\circ$ , 194-195 $^\circ$ , pK $^{25}$  7.80 (phenolic OH)**. Purify it by recrystallisation from EtOH. It is very slightly soluble in cold  $\text{H}_2\text{O}$  (solubility at  $37^\circ$  is 0.22%), slightly soluble in  $\text{Et}_2\text{O}$  and  $\text{CHCl}_3$ , but soluble in MeOH and

AcOH. It has a blue fluorescence in aqueous EtOH and has UV with  $\lambda_{\text{max}}$  at 221, 251 and 322.5nm (MeOH). The IR has  $\nu_{\text{max}}$  at 3077 br, 1667, 1592, 1385, 1267, 1156, 1130 and 1066  $\text{cm}^{-1}$ . The *acetate* has **m** 153-154°. [Woods & Sapp *J Org Chem* **27** 3703 1962, *Beilstein* **18** III/IV 332, **18/1** V 439.]

**2-Hydroxymethyl-12-crown-4** [75507-26-5] **M 206.2, b 115°/0.04mm, d**<sub>4</sub><sup>20</sup> **1.186, n**<sub>D</sub><sup>20</sup> **1.480**. Purify it by chromatography on Al<sub>2</sub>O<sub>3</sub> with EtOAc as eluent to give a *hygroscopic* colourless oil, which is distilled under high vacuum and is distilled *in vacuo*. It has IR with  $\nu_{\text{max}}$  at 3418 (OH) and 1103 (COC)  $\text{cm}^{-1}$ , and NMR:  $\delta$  3.70s. [Kimura et al. *J Chem Soc, Chem Commun* 492 1983, Pugia et al. *J Org Chem* **52** 2617 1987.]

**S-(-)-5-Hydroxymethyl-2(5H)-furanone** [78508-96-0] **M 114.1, 39-42°, 40-44°, b 130°/0.3mm, [** $\alpha$ <sub>D</sub><sup>20</sup> **-180°, [** $\alpha$ <sub>D</sub><sup>20</sup> **-148° (c 1.4, H<sub>2</sub>O)**. It is purified by chromatography on Silica gel using hexane/EtOAc (1:1) to give a colourless oil which is distilled using a Kügelrohr apparatus, and the distillate crystallises on cooling. It has R<sub>F</sub> 0.51 on Whatman No 1 paper using pentan-1-ol and 85% formic acid (1:1) and developing with ammoniacal AgNO<sub>3</sub>. [Boll *Acta Chem Scand* **22** 3245 1968, NMR: Oppolzer et al. *Helv Chim Acta* **68** 2100 1985, *Beilstein* **18** III/IV 56, **18/1** V 54.]

**5-(Hydroxymethyl)furfural** [67-47-0] **M 126.1, m 33.5°, b 114-116°/1mm, d**<sub>25</sub><sup>25</sup> **1.2620, n**<sub>D</sub><sup>25</sup> **1.5533**. Crystallise it from diethyl ether/petroleum ether. [*Beilstein* **18** III/IV 100, **18/1** V 130.]

**dl-3-Hydroxy-N-methylmorphinan** [297-90-5] **M 257.4, m 251-253°**. Crystallise it from MeOH or aqueous EtOH. The *hydrochloride* has **m** 176-177°. The *hydrobromide* has **m** 200-201° when recrystallised from EtOH/Et<sub>2</sub>O. [Sneider & Hellerbach *Helv Chim Acta* **33** 1446 1950, *Beilstein* **21** III/IV 100.]

**8-Hydroxy-2-methylquinoline** [826-81-3] **M 159.2, m 74-75°, b 266-267°, pK**<sub>1</sub><sup>25</sup> **5.61, pK**<sub>2</sub><sup>25</sup> **10.16**. Crystallise the quinoline from EtOH or aqueous EtOH. Its solubility at 20° in H<sub>2</sub>O is 0.366g/L, and in CHCl<sub>3</sub> it is 466g/L. It complexes with many metals. [*Beilstein* **21** H 106, **21** III/IV 2132.]

**4-Hydroxy-2-n-nonylquinoline N-oxide** [316-66-5] **M 287.4, m 148-149°, pK**<sub>Est</sub> **~6.0**. Crystallise the *N*-oxide from EtOH. Its UV has  $\lambda_{\text{max}}$  at 220-230nm in 0.001N NaOH. [Cornforth & James *Biochem J* **63** 124 1956, *Beilstein* **21** III/IV 3834.]

**1-Hydroxyphenazine (Hemipyocyanine)** [528-71-2] **M 196.2, m 157-158°, 165-166°, pK**<sub>1</sub><sup>15</sup> **1.61, pK**<sub>2</sub><sup>15</sup> **8.33 (10% aqueous MeOH)**. Purify the hydroxyphenazine by chromatography on acidic alumina with \*benzene/ether, and recrystallise it from aqueous EtOH, \*benzene/heptane, then sublime it at 140°/1.5mm. [UV, IR: Badger et al. *J Chem Soc* 3204 1951, Hegedüs *Helv Chim Acta* **33** 746 1950, *Beilstein* **23** II 360, **23** III/IV 2753.]

**2-(2-Hydroxyphenyl)benzothiazole** [3411-95-8] **M 227.2, m 132-133°, b 173-179°/3mm**. Recrystallise it several times from aqueous EtOH or dilute AcOH and sublime it. [Itoh & Fujiwara *J Am Chem Soc* **107** 1561 1985, Bogert & Corbitt *J Am Chem Soc* **48** 786 1926, *Beilstein* **27** II 91.]

**2-(2-Hydroxyphenyl)benzoxazole** [835-64-3] **M 211.2, m 127°, b 338°/760mm**. Recrystallise it several times from aqueous EtOH or dilute AcOH and sublime it. An aqueous alkaline solution containing EtOH has a blue fluorescence. [Itoh & Fujiwara *J Am Chem Soc* **107** 1561 1985, *Beilstein* **27** II 91.]

**3-Hydroxy-2-phenylcinchoninic acid (Oxycincophen)** [485-89-2] **M 265.3, m 206-207°(dec)**. It is precipitated from alkaline solution on acidification and crystallises from EtOH or AcOH in yellow crystals. [Marshall & Blanchard *J Pharmacol* **95** 186 1949, *Beilstein* **22** H 245, **22** II 183, **22** III/IV 2383.]

**(±)-2-( $\alpha$ -Hydroxypropyl)piperidine [2-piperidinepropanol, *RS*-1-(*RS*-2-piperidyl)propan-1-ol, ( $\pm$ )- $\alpha$ -conhydrine] [3238-62-8, 63401-12-7, 24448-89-3] **M 143.2, m 99-100°, pK**<sub>Est</sub> **~10.2**. Crystallise it from ether. The [*RS*-1-(*SR*-2)]-isomer also crystallises from ether and has **m** 98-98.5°. The *methiodide* crystallises from Me<sub>2</sub>CO/MeOH with **m** 127-130.5°. **POISONOUS** [Sicher & Tichy *Col Czech Chem Commun* **23** 2081 1958, Govindachari & Rajappa *J Chem Soc* 1306 1958, Stereochemistry: Hill *J Am Chem Soc* **80** 1609 1958,**

*Beilstein* 21 II 21, 21 III/IV 122.]

(+)-2-( $\alpha$ -Hydroxypropyl)piperidine [2-piperidinepropanol, *R*-1-(*S*-2-piperidyl)propan-1-ol, (+)- $\alpha$ -conhydrine] [495-20-5] M 143.2, m 121°, b 224-5°/720mm,  $[\alpha]_D^{20}$  -9.8° (c 4, EtOH), pK<sub>Est</sub> ~10.2. This very POISONOUS alkaloid from hemlock crystallises in leaflets from ether. The *O,N*-dibenzoyl derivative has m 133-134° and  $[\alpha]_D^{24}$  -13° (c 3, CHCl<sub>3</sub>). [Sicher & Tichy *Col Czech Chem Commun* 23 2081 1958, Stereochemistry: Hill *J Am Chem Soc* 80 1609 1958, Absolute config & ORD: Fodor et al. *Canad J Chem* 47 4393 1969, *Beilstein* 21 I 191, 21 II 21, 21 III/IV 122.]

(-)-2-( $\alpha$ -Hydroxypropyl)piperidine [2-piperidinepropanol, *R*-1-(*R*-2-piperidyl)propan-1-ol, (-)- $\alpha$ -conhydrine] [495-20-5] M 143.2, m 121°, b 224-5°/720mm,  $[\alpha]_D^{20}$  +10° (c 10, EtOH), pK<sub>Est</sub> ~10.2. Crystallise the piperidine from ether. POISONOUS. [Galinovskiy & Mulley *Monatsh Chem* 79 426 1948, *Beilstein* 21 I 191, 21 II 21, 21 III/IV 122.]

( $\pm$ )-7-(2-Hydroxypropyl)theophylline (Proxophylline, 1,3-dimethyl-3*H*,7*H*-purine-2,6-dione) [603-00-9] M 238.2, m 135-136°. Crystallise it from EtOH, aqueous MeOH or EtOAc. Roth *Archiv Pharmazie* 292 234 1959, Zelnik et al. *Bull Soc Chim Fr* 1733 1956, *Beilstein* 26 III/IV 2366.]

6-Hydroxypurine (hypoxanthine) [68-94-0] M 136.1, m 150°(dec), pK<sub>1</sub><sup>20</sup> 1.98, pK<sub>2</sub><sup>20</sup> 8.96, pK<sub>3</sub><sup>20</sup> 12.18. Crystallise it from hot water and dry it at 105°. [*Beilstein* 26 II 252, 26 III/IV 2081.]

2-Hydroxypyridine (2-pyridone) [142-08-5] M 95.1, m 105-107°, b 181-185°/24mm,  $\epsilon_{293\text{nm}}$  5900 (H<sub>2</sub>O) pK<sub>1</sub><sup>25</sup> 1.25, pK<sub>2</sub><sup>25</sup> 11.99. Distil the pyridone under vacuum to remove coloured impurity, then recrystallise from \*benzene, CCl<sub>4</sub>, EtOH or CHCl<sub>3</sub>/diethyl ether. It can be sublimed under high vacuum. [DePue et al. *J Am Chem Soc* 107 2131 1985, *Beilstein* 21/7 V 106.]

3-Hydroxypyridine [109-00-2] M 95.1, m 129°, 130°, pK<sub>1</sub><sup>25</sup> 5.10, pK<sub>2</sub><sup>25</sup> 8.6. Crystallise 3-hydroxypyridine from \*C<sub>6</sub>H<sub>6</sub>, water or EtOH. [*Beilstein* 21 III/IV 402, 21/2 V 68.]

4-Hydroxypyridine (4-pyridone) [626-64-2] M 95.1, m 65°, 68° (hydrate), 148.5°, 151-152° (anhydr), b >350°/760mm, pK<sub>1</sub><sup>20</sup> 3.20, pK<sub>2</sub><sup>20</sup> 11.12. Crystallise 4-pyridone from H<sub>2</sub>O or wet CHCl<sub>3</sub> as the *monohydrate*. It loses H<sub>2</sub>O on drying *in vacuo* over H<sub>2</sub>SO<sub>4</sub>. Store it over KOH because it is *hygroscopic*. [*Beilstein* 21 III/IV 446, 21/7 V 152.]

2(6)-Hydroxypyridine-5(3)-carboxylic acid (6-hydroxynicotinic acid) [5006-66-6] M 139.1, m 304°(dec), pK<sub>1</sub><sup>20</sup> 3.82 and pK<sub>2</sub><sup>20</sup> 9.92. It crystallises from water with m 303.4-303.7°(dec), or with m 325°(dec) from aqueous EtOH. The *methyl ester* crystallises from Me<sub>2</sub>CO with m 166° and pK<sup>20</sup> 9.92. [Albert *J Chem Soc* 1020 1960, *Beilstein* 22 III/IV 2147, 22/6 V 119.]

4-Hydroxypyridine-2,6-dicarboxylic acid (chelidamic acid) [134.828-60-3] M 183.1, m 248°(dec, 1 H<sub>2</sub>O), 254°(dec), 263°(dec), pK<sub>1</sub><sup>22</sup> 1.9, pK<sub>2</sub><sup>22</sup> 3.18, pK<sub>3</sub><sup>22</sup> 10.85; pK<sub>1</sub> 3.62, pK<sub>2</sub> 4.82 (80% aqueous methoxyethanol). It crystallises from water. The *dimethyl ester* crystallises from MeOH with m 167° (*monohydrate*, from H<sub>2</sub>O), m 170-171° (*anhydrous*, from MeOH), and pK<sup>25</sup> 6.25. [Bensaude et al. *J Am Chem Soc* 99 4438 1972, *Beilstein* 22 III/IV 2583, 22/7 V 164.]

2-Hydroxypyrimidine [557-01-7] M 96.1, m 179-180°, pK<sub>1</sub><sup>20</sup> 2.24, pK<sub>2</sub><sup>20</sup> 9.17. It crystallises from EtOH or ethyl acetate. Its solubility in H<sub>2</sub>O at 20° is 1g/2.2ml. [Albert *J Chem Soc* 4219 1952, *Beilstein* 24 III/IV 194.], 2-Hydroxypyrimidine hydrochloride [38353-09-2] M 132.5, has m 203-205.5°(dec) and crystallises from EtOH. [*Beilstein* 24 III/IV 173.] The *picrate* has m 199° and crystallises from EtOH. [Brown et al. *J Chem Soc* 211 1955.]

4-Hydroxypyrimidine [4562-27-0] M 96.1, m 164-165°, pK<sub>1</sub><sup>20</sup> 1.66, pK<sub>2</sub><sup>20</sup> 8.63. It crystallises from \*benzene or ethyl acetate. The *picrate* has m 164-166° and crystallises from EtOH [Brown et al. *J Chem Soc* 4035 1955, *Beilstein* 24 III/IV 171.]

**R-(+)-3-Hydroxypyrrolidine** [ $(\pm)$  2799-21-5 and 104706-47-0,  $R(+)$  40499-83-0,  $S(-)$  100243-39-8] **M 87.1**, **b 215-216°/atm**,  $d_4^{20}$  **1.078**,  $n_D^{20}$  **1.490**,  $[\alpha]_D^{20}$  **+6.5°** (c 1.5, MeOH),  $pK_{Est}$  **~10.1**. The  $(\pm)$ -isomer is purified by repeated distillation (**b** 102-104°/12mm, 108-110°/18mm), and the  $(\pm)$ -picrate crystallises from EtOH with **m** 140-141°. The  $R(+)$ -enantiomer has **b** 70°/0.6mm and  $[\alpha]_D^{20}$  **+5.6°** (c 3.63, MeOH). Its *hydrochloride* has a negative rotation and its *dimethiodide* has **m** 230° and  $[\alpha]_D^{24}$  **-8.02°**. [Suyama & Kanno *Yakugaku Zasshi (J Pharm Soc Japan)* **85** 531 1965, Uno et al. *J Heterocycl Chem* **24** 1025 1987, Flanagan & Joullie *Heterocycles* **26** 2247 1987, *Beilstein* **21** III/IV 44.]

**2-Hydroxyquinoline (carbostyryl)** [59-31-4] **M 145.2**, **m 199-200°**,  $pK_1^{20}$  **-0.31**,  $pK_2^{20}$  **11.76**. Crystallise it from MeOH. It has **m** 200-201° after sublimation in a vacuum. The *picrate* has **m** 132° after crystallisation from Et<sub>2</sub>O. [Gibson et al. *J Chem Soc* 4340 1955, *Beilstein* **21** III/IV 1057, **21/8** V 217.]

**8-Hydroxyquinoline (oxine, 8-quinolinol)** [148-24-3] **M 145.2**, **m 71-73°, 75-76°, 76°**, **b 122°/0.1mm**, **~267°/atm**,  $pK_1^{25}$  **4.91**,  $pK_2^{25}$  **9.81**. Crystallise oxine from hot EtOH, acetone, petroleum ether (b 60-80°) or water. Crude oxine can be purified by precipitation of copper oxinate, followed by liberation of free oxine with H<sub>2</sub>S or by steam distillation after acidification with H<sub>2</sub>SO<sub>4</sub>. Store it in the dark. It forms complexes with many metals. [Manske et al. *Can J Research* **27F** 359 1949, Phillips *Chem Rev* **56** 271 1956, *Beilstein* **21** III/IV 1135, **21/3** V 252.]

**8-Hydroxyquinoline-5-sulfonic acid (H<sub>2</sub>O)** [84-88-8] **M 243.3**, **m 322-323°(sintering at ~305°), >310°**,  $pK_1^{25}$  **4.09**,  $pK_2^{25}$  **8.66**. Crystallise the acid from water (as the 1.5 *hydrate*, **m** 316-317°) or dilute HCl (ca 2% by weight). [*Beilstein* **22** I 620, **22** II 313, **22** III/IV 3493.]

**4-Hydroxy-2,2,6,6-tetramethylpiperidine** [2403-88-5] **M 157.3**, **m 130-131°**,  $pK^{20}$  **10.05**. The piperidine crystallises from water as a *hydrate* and crystallises from dry ether or \*C<sub>6</sub>H<sub>6</sub> as the *anhydrous* base. The *hydrochloride* has **m** 282-284° (from EtOH/H<sub>2</sub>O), and the *formate* has **m** 207°(dec, from EtOH/EtOAc). [Mailey & Day *J Org Chem* **22** 1061 1957, *Beilstein* **21** I 195, **21** III/IV 146, **21/1** V 159.]

**4(6)-Hydroxy-2,5,6(2,4,5)-triaminopyrimidine sulfate** [35011-47-3] **M 257.22**, **m >340°**,  $pK_1$  **2.0**,  $pK_2$  **5.1**,  $pK_3$  **10.1**. This salt has very low solubility in H<sub>2</sub>O. It is best purified by conversion into the dihydrochloride salt, which is then re-converted to the insoluble sulfate salt. The sulfate salt (2.57g, 10mmoles) is suspended in H<sub>2</sub>O (20ml) containing BaCl<sub>2</sub> (10mmoles) and stirred in a boiling water bath for 15 minutes. After cooling, the insoluble BaSO<sub>4</sub> is filtered off and washed with boiling H<sub>2</sub>O (10ml). The combined filtrate and washings are made acidic with HCl and evaporated to dryness. The residual hydrochloride salt is recrystallised from H<sub>2</sub>O by adding conc HCl whereby the *dihydrochloride salt* separates as clusters which darken at 260° and dec > 300°[darkening > 360°]. [Baugh & Shaw *J Org Chem* **29** 3610 1964, King & Spengley *J Chem Soc* 2144 1952]. The hydrochloride is then dissolved in H<sub>2</sub>O, and while hot an equivalent of H<sub>2</sub>SO<sub>4</sub> is added when the sulfate separates as a white microcrystalline solid which is filtered off washed liberally with H<sub>2</sub>O and dried in vacuum over P<sub>2</sub>O<sub>5</sub>. [Albert & Wood *J Appl Chem London* **3** 521 1953, UV: Cavalieri et al. *J Am Chem Soc* **70** 3875 1948; see also Pfeleiderer *Chem Ber* **90** 2272 1957, Traube *Chem Ber* **33** 1371 1900.] The *hydrochloride* has **m** > 300°. [*Beilstein* **25** II 384, **25** III/IV 3648.]

**3-Hydroxyxanthone** [3722-51-8] **M 212.2**, **m 246°, 249-250°**. Purify the xanthone by chromatography on SiO<sub>2</sub> gel with petroleum ether/\*benzene as eluent. Recrystallise it from \*benzene or EtOH. The *acetate* has **m** 157-158°. [Itoh et al. *J Am Chem Soc* **107** 4819 1985, Davies et al. *J Org Chem* **23** 307 1958]. [*Beilstein* **18** H 46, **18** I 315, **18** II 29, **21** III/IV 601.]

**Ibogaine (12-methoxybogamine)** [83-74-9] **M 300.3**, **m 152-153°**,  $[\alpha]_D^{20}$  **-54°** (EtOH),  $pK^{25}$  **8.1** (**80% aqueous MeOCH<sub>2</sub>CH<sub>2</sub>OH**). Crystallise this alkaloid from EtOH or aqueous EtOH and sublime it at 150°/0.01mm. It is soluble in organic solvents but insoluble in H<sub>2</sub>O. The *hydrochloride*, **m** 299-300°(dec), is soluble in H<sub>2</sub>O and alcohols. [Büchi et al. *J Am Chem Soc* **88** 3099 1866, Rosenmund *Chem Ber* **108** 1871 1975, *Beilstein* **23** III/IV 2742.]

**Imidazole (glyoxaline)** [288-32-4] **M 68.1, m 89.5-91°, 89-90°, b 140-145°/15mm, 256°/atm, 262-264°/atm, pK<sub>1</sub><sup>25</sup> 6.99, pK<sub>2</sub><sup>25</sup> 14.44.** Crystallise imidazole from \*benzene, CCl<sub>4</sub>, CH<sub>2</sub>Cl<sub>2</sub>, EtOH, petroleum ether, acetone/petroleum ether and distilled de-ionized water. Dry it at 40° under vacuum over P<sub>2</sub>O<sub>5</sub>. Distil it at low pressure. It is also purified by sublimation or by zone melting. [Snyder et al. *Org Synth Coll Vol III* 471 1955, Bredereck et al. *Chem Ber* **97** 827 1964, Caswell & Spiro *J Am Chem Soc* **108** 6470 1986.] <sup>15</sup>N-imidazole crystallises from \*benzene [Scholes et al. *J Am Chem Soc* **108** 1660 1986]. [*Beilstein* **23** II 34, **23** III/IV 564, **23/4** V 191.]

**1H-Indazole-3-carboxylic acid** [4498-67-3] **M 162.2, m 265-265.5°, 268-268.5°, pK<sub>Est(1)</sub> ~4.5.** Purify the acid by recrystallisation from glacial acetic acid (charcoal), and dry the yellow crystals *in vacuo*. Alternatively, dissolve the acid in boiling H<sub>2</sub>O, concentrate the solution to one third its volume, cool and collect the yellow powder. Its UV has λ<sub>max</sub> (H<sub>2</sub>O) at 295 (log<sub>e</sub> 3.88) nm. The *ethyl ester* (prepared via the acid chloride) has **m 139°** (from EtOH) and has λ<sub>max</sub> (MeOH) at 393 nm, and the *N-methylamide* crystallises in tan needles from MeNO<sub>2</sub> with **m 191-191° (187-188.5°** was also reported). [Rousseau & Lindwall *J Am Chem Soc* **72** 3048 1950, Snyder et al. *J Am Chem Soc* **74** 2009 1952, *Beilstein* **25** H 129, **25** I 238, **25** II 128, **25** IV 808.]

**1H, 2H-Indazol-3-one (3-hydroxy-1H-indazole)** [7364-25-2] **M 134.1, m 250-252°, 253-254°, pK<sub>Est(1)</sub> ~6.0.** Purify indazol-3-one by recrystallisation from MeNO<sub>2</sub> or aqueous MeOH, and sublimation at 220°/0.1mm. In the UV it has λ<sub>max</sub> in MeOH at 215 and 306 nm, and it is in the 3-OH form in ethanolic solution. There appears to be some controversy regarding the tautomeric form in the solid state, but the IR peak at ν<sub>max</sub> 1627 cm<sup>-1</sup> (KBr) supports the keto form. [Ainsworth *J Am Chem Soc* **79** 524 1957, Elguero et al. *Adv Heterocycl Chem, Suppl* **1** 354 1976, Chernokal'skii et al. *Khim Geterotsikl Soedin* **96** 1966, *Beilstein* **24** H 111, **24** II 59, **24** III/IV 270.]

**4'-(Imidazol-1-yl)acetophenone** [10041-06-2] **M 186.2, m 104-107°, pK<sup>25</sup> 4.54.** Recrystallise it twice from CH<sub>2</sub>Cl<sub>2</sub>/hexane [Collman et al. *J Am Chem Soc* **108** 2588 1986].

**2-Iminothiolane hydrochloride (2-iminotetrahydrothiophene, Traut's reagent)** [4781-83-3] **M 137.6, m 187-192°, 190-195°, 192-193°, 193-194°, 202-203°, pK<sup>25</sup> <2 (free base).** Recrystallise the hydrochloride from MeOH/Et<sub>2</sub>O (**m 187-192°**) or (MeOH/Me<sub>2</sub>CO), but after sublimation at ~180°/0.2mm the melting point rises to 202-203°. It has <sup>1</sup>HNMR with δ 2.27 (2H, t), 3.25 (2H, t) and 3.52 (2H, t) in -(CD<sub>3</sub>)<sub>2</sub>SO. [King et al. *Biochemistry* **17** 1499 1978.] The *free base* is purified by vacuum distillation (**b 71-72°/6mm**), has IR (film) with ν<sub>max</sub> at 1700 (C=N)cm<sup>-1</sup> and <sup>1</sup>HNMR (CDCl<sub>3</sub>) with δ at 3.58 (2H, t) and 2.10-2.8 (4H, m). The *free base* is stable on storage but slowly hydrolyses in aqueous solutions with half-lives at 25° of 390 hours at pH 9.1, 210 hours at pH 10 and 18 hours at pH 11. [Traut et al. *Biochemistry* **12** 3266 1973, *Biochemistry* **17** 399 1978, Alagon & King *Biochemistry* **19** 4343 1980, *Beilstein* **17/9** V 12.]

**Indanthrene (N,N'-dihydro-1,2,1',2'-anthraquinonazine)** [81-77-6] **M 442.4, m 470-500°(dec).** Crystallise indanthrene repeatedly from 1,2,4-trichlorobenzene to give a blue powder. It is soluble in alkali and conc H<sub>2</sub>SO<sub>4</sub>. [Weinstein & Merritt *J Am Chem Soc* **81** 3759 1959, *Beilstein* **24** II 317, **24** III/IV 2193.]

**Indazole** [271-44-3] **M 118.1, m 147°, 150°, pK<sub>1</sub><sup>20</sup> 1.32, pK<sub>2</sub><sup>20</sup> 13.80 (acidic NH).** Crystallise indazole from water, sublime it *in vacuo*, then recrystallise it from petroleum ether (**b 60-80°**). The *picrate* crystallises from Et<sub>2</sub>O with **m 136°**. [Ainsworth *Org Synth Coll Vol IV* 536 1963, *Beilstein* **23** III/IV 1055, **23/6** V 156.]

**Indigo** [482-89-3] **M 262.3, sublimes at ~300°, m 390°(dec), and halogen-substituted indigo dyes.** First reduce indigo in alkaline solution with sodium hydrosulfite, and filter. The filtrate is then oxidised by air, and the resulting precipitate is filtered off, dried at 65-70°, ground to a fine powder, and extracted with CHCl<sub>3</sub> in a Soxhlet extractor. Evaporation of the CHCl<sub>3</sub> extract gives the purified dye. [Brode et al. *J Am Chem Soc* **76** 1034 1954; spectral characteristics are listed, *Beilstein* **24** II 233, **24** III/IV 1791.]

**Indole** [120-72-9] **M 117.2, m 52°, 54.5°, b 124°/5mm, 253-254°/760mm, pK<sub>1</sub><sup>25</sup> -2.47 (H<sub>0</sub> scale), pK<sub>2</sub><sup>25</sup> 16.97 (acidic NH).** Crystallise indole from \*benzene, hexane, petroleum ether, water or EtOH/water (1:10). It can be further purified by sublimation in a vacuum or by zone melting. The *picrate* forms orange crystals from

EtOH and has  $m$  175°. [*Beilstein* 20 II 196, 20 III/IV 3176, 20/7 V 5.]

**Indole-3-acetic acid (heteroauxin)** [87-51-4]  $M$  175.2,  $m$  167-169°(dec),  $pK_i^{25}$  -6.13 (aqueous  $H_2SO_4$ ),  $pK_2^{25}$  4.54 ( $CO_2H$ ). Recrystallise heteroauxin from EtOH/water [James & Ware *J Phys Chem* 89 5450 1985]. [*Beilstein* 22 III/IV 65.] Alternatively, recrystallise 30g of the acid with 10g of charcoal in 1L of hot water, filter and cool when 22g of colourless acid separate. Dry it and store it in a dark bottle away from direct sunlight [Johnson & Jacoby *Org Synth Coll Vol V* 654 1973]. The *picrate* has  $m$  178-180°. [*Beilstein* 22 H 66, 22 I 508, 22 II 50, 22 III/IV 1088.] It is a plant growth substance.

**3-Indoleacetonitrile** [771-51-7]  $M$  156.2,  $m$  33-36°, 36-38°,  $b$  157°/0.2mm, 158-160°/0.1mm, viscous oil  $n_D^{20}$  1.6097. Distil the nitrile at very high vacuum, and the viscous distillate crystallises on standing after a few days; the *picrate* has  $m$  127-128° (from EtOH) [Coker et al. *J Org Chem* 27 850 1962, Thesing & Schülde *Chem Ber* 85 324 1952]. Store it away from light. The *N-acetate* has  $m$  118° (from MeOH) and has  $R_F = 0.8$ , on Silica Gel F<sub>254</sub> in  $CHCl_2/MeOH$  19:1 [Buzas et al. *Synthesis* 129 1977]. [*Beilstein* 22 III/IV 1097, 22/3 V 74.] It is a plant growth substance.

**Indole-2-carboxaldehyde** [19005-93-7]  $M$  145.2,  $m$  138°, 140-141°, 138-142°,  $pK_{Est} < 1$ . Recrystallise the indole from aqueous MeOH ( $m$  141-142°),  $Et_2O$  ( $m$  138°) or sublime it *in vacuo* ( $m$  138°). The *N,N*-dimethylhydrazone [127280-17-5] has  $m$  105-106° (from hexane/\* $C_6H_6$ ). The *thiosemicarbazone* has  $m$  229° (dec) (from 50% aqueous EtOH) [Doyle et al. *J Chem Soc* 2854 1956]. The *2,4-dinitrophenylhydrazone* has  $m$  315-320° (dec) (from pyridine/MeOH). [Suzuki et al. *Chem Pharm Bull Jpn* 39 2170 1991, *Beilstein* 21 III/IV 3754.]

**Indole-3-propionic acid** [830-96-6]  $M$  189.2,  $m$  134-135°,  $pK^{25}$  4.95. Recrystallise it from EtOH/water [James & Ware *J Phys Chem* 89 5450 1985]. The *picrate* has  $m$  143-144°, and the *methyl ester* crystallises from \* $C_6H_6$  or MeOH with  $m$  81-83°. [*Beilstein* 22 III/IV 1113, 22/3 V 114.]

(±)-**Indoline-2-carboxylic acid (2,3-dihydro-1H-indole-2-carboxylic acid)** [78348-24-0]  $M$  163.2,  $m$  168°(dec),  $pK_{Est(1)} \sim 2.1$ ,  $pK_{Est(2)} \sim 3.8$ . Dissolve the acid in hot EtOH, add excess of dry  $Et_2O$  and cool to yield colourless plates that decompose in the range of 120-150°. The *amide* ( $m$  208-209°) crystallises as colourless plates which sublime at 150°/1.0mm and has  $\nu_{max}$  at 1625 $cm^{-1}$  (paraffin mull) [Hudson & Robertson *Aust J Chem* 20 1935 1967]. [*Beilstein* 22/2 V 421.] See Chapter 6, Catalysts-Part 1 for optical isomers

**Indolizine [pyrrocoline, pyrrolo(1,2-a)pyridine]** [274-40-8]  $M$  117,  $m$  73-74°, 75°,  $pK^{20}$  3.94 (**C-protonation**). Purify indolizine through an alumina column in \* $C_6H_6$  and elute with \* $C_6H_6$  (toluene could be used instead). The eluate contained in the fluorescent band (using UV light  $\lambda$  365nm) is collected, evaporated and the crystalline residue is sublimed twice at 40-50°/0.2-0.5mm. The colourless crystals have a *naphthalene* odour, darken on standing and should be stored in dark sealed containers. If the original sample is dark in color then it should be covered with water and steam distilled. The colourless crystals in the distillate are collected and dried between filter paper and sublimed. It protonates on C3 in aqueous acid. It should give one fluorescent spot on paper chromatography (Whatman 1) in 3% aqueous ammonia and in *n*-BuOH/AcOH/ $H_2O$  (4:1:1). The *picrate* has  $m$  101° from EtOH. [Armarego *J Chem Soc* 226 1964, Armarego *J Chem Soc (B)* 191 1966, Scholtz *Chem Ber* 45 734 1912, *Beilstein* 20 II 200, 20 III/IV 3195.]

**trans-Indol-3-ylacrylic acid** [1204-06-4]  $M$  187.2,  $m$  190-195°(dec), 195°(dec), 196°(dec), 195-196°(dec),  $pK_{Est} \sim 4.2$ . Recrystallise the acid from AcOH,  $H_2O$  or EtOAc/cyclohexane. UV in MeOH has  $\lambda_{max}$  at 225, 274 and 325nm. [Shaw et al. *J Org Chem* 23 1171 1958, constitution: Rappe *Acta Chem Scand* 18 818 1964, Moffatt *J Chem Soc* 1442 1957, Kimming et al. *Hoppe Seyler's Z Physiol Chem* 371 234 1958, *Beilstein* 22 V 249.]

**3-Indolylbutyric acid** [133-32-4]  $M$  203.2,  $m$  120-123°, 123-125°, 124°,  $pK^{25}$  4.84. Recrystallise the acid from  $H_2O$ . It is soluble in EtOH,  $Et_2O$  and  $Me_2CO$  but insoluble in  $CHCl_3$ . [Bowman & Islip *Chem Ind London* 154 1971, Jackson & Manske *J Am Chem Soc* 52 5029 1930, Albaum & Kaiser *Am J Bot* 24 420 1937.] It has also been recrystallised from EtOH/water [James & Ware *J Phys Chem* 89 5450 1985]. Its UV

has  $\lambda_{\max}$  at 278 and 320nm in isoPrOH [Elvidge *Quart J Pharm Pharmacol* **13** 219 1940]. The *methyl ester* has **m** 73-74° (from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether) and **b** 230°/6mm [Bullock & Hand *J Am Chem Soc* **78** 5854 1951]. [Beilstein **22** III/IV 1128, **22/3** V 140.]

**3-Indolylpyruvic acid** [392-12-1] **M** 203.2, **m** ~210°(dec), **208-210°(dec)**, **219°(dec)**, **pK<sub>Est</sub>** ~2.4. Recrystallise the acid from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>, EtOAc/CHCl<sub>3</sub>, Me<sub>2</sub>CO/AcOH (crystals have 1 molecule of AcOH) and dioxane/\*C<sub>6</sub>H<sub>6</sub> (with 0.5 molecule of dioxane) [Shaw et al. *J Org Chem* **23** 1171 1958, Kaper & Veldstra *Biochim Biophys Acta* **30** 401 1958]. The *ethyl ester* has **m** 133° (from Et<sub>2</sub>O), and its *2,4-dinitrophenylhydrazone* has **m** 255° (from Me<sub>2</sub>CO). [Baker *J Chem Soc* 461 1946.] The *oxime* has **m** 157°(dec, from EtOAc/Et<sub>2</sub>O) and **pK<sup>20</sup>** 3.40 [Ahmad & Spenser *Canad J Chem* **39** 1340 1961]. [Beilstein **22** II 250, **22** III/IV 3080, **22/6** V 324.]

**7-Iodoindole** [89976-15-8] **M** 243.0, **m** 52-56°, **pK<sub>Est</sub>** <1. Purify 7-iodoindole by chromatography through a silica gel column and eluting with CH<sub>2</sub>Cl<sub>2</sub>/hexane (1:3, v/v) followed by recrystallisation from hexane (colourless plates, **m** 55-56°). [Somei et al. *Chem Pharm Bull Jpn* **35** 3146 1987, Somei & Saida *Heterocycles* **23** 3114 1985.]

**Iodinine (1,6-dihydroxyphenazine-5,10-dioxide)** [68-81-5] **M** 244.1, **m** 236°(dec), **pK<sup>25</sup>** 12.5. Purify iodinine through a column of silica gel and elute with Me<sub>2</sub>CO/CHCl<sub>3</sub>, then recrystallise it from CHCl<sub>3</sub> to give purple crystals with a copper-coloured luster. [Clemons & Dalgleish *J Chem Soc* 1481 1950, Gerber & Lechevalier *Biochemistry* **3** 598 1964, Beilstein **23** III/IV 3227,]

**Iodonitrotetrazolium chloride (2[4-iodophenyl]-3-[4-nitrophenyl]-5-phenyl-2H-tetrazolium chloride)** [146-68-9] **M** 505.7, **m** 229°(dec), ~245°(dec). Recrystallise the chloride from H<sub>2</sub>O, aqueous EtOH or EtOH/Et<sub>2</sub>O. Alternatively, dissolve it in the minimum volume of EtOH and add Et<sub>2</sub>O; or dissolve it in hot H<sub>2</sub>O (charcoal), filter and precipitate it by adding conc HCl. Filter the solid off and dry it at 100°. Its solubility in H<sub>2</sub>O at 25° is 0.5%, and in hot MeOH/H<sub>2</sub>O (1:1) it is 5%. [Fox & Atkinson *J Am Chem Soc* **72** 3629 1950, Beilstein **26** III/IV 1776.]

**Iodonitrotetrazolium violet-Formazan** [7781-49-9] **M** 471.3, **m** 185-186°. Dissolve it in boiling dioxane (20g in 300ml), add H<sub>2</sub>O (100ml) slowly, cool, filter and dry it *in vacuo* at 100°. Its solubility in CHCl<sub>3</sub> is ~1%. [UV: Fox & Atkinson *J Am Chem Soc* **72** 3629 1950, Beilstein **26** III/IV 1776.]

**Isatin (indole-2,3-dione)** [91-56-5] **M** 147.1, **m** 201-203°, 205°, **pK<sup>25</sup>** >12 (acidic NH). Crystallise isatin from amyl alcohol and sublime it at 180°/1mm. In aqueous NaOH the ring opens to yield sodium *o*-aminobenzoylformate. [Beilstein **21** II 327, 567, **21** III/IV 4981, **21/10** V 221.]

**Isatoic anhydride (3,1-benzoxazin-2,4[1-H]-dione)** [118-48-9] **M** 163.1, **m** 235-240°, 240-243°, 243°, 243-245°. Recrystallise it from EtOH or 95% EtOH (30ml/g) or dioxane (10ml/g) and dry it in a vacuum. [Wagner & Fegley *Org Synth Coll Vol III* 488 1955, Ben-Ishai & Katchalski *J Am Chem Soc* **74** 3688 1952, UV: Zentmyer & Wagner *J Org Chem* **14** 967 1949, Beilstein **27** II 299, **27** III/IV 3330.]

**3-Isobutyl-1-methylxanthine (3-isobutyl-1-methylpurine-2,6-dione)** [28822-58-4] **M** 222.3, **m** 199-210°, 202-203°, **pK<sub>Est</sub>** ~ 6.7 (acidic NH). Recrystallise it from aqueous EtOH. [Beilstein **26** III/IV 2350.]

**(+)-Isolysergic acid** [478-95-5] **M** 268.3, **m** 218°(dec),  $[\alpha]_{\text{D}}^{20}$  +281° (c 1, pyridine) **pK<sub>1</sub><sup>24</sup>** 3.33, **pK<sub>2</sub><sup>24</sup>** 8.46. It crystallises from water as the *dihydrate*. The *methyl ester* has **m** 172-174° (from MeOH or \*C<sub>6</sub>H<sub>6</sub>) and  $[\alpha]_{\text{D}}^{20}$  +179° (c 0.5 CHCl<sub>3</sub>). [Smith & Timmis *J Chem Soc* 1440 1936, Craig et al. *J Biochem* **125** 289 1938, Stenlake *J Chem Soc* 1626 1955, Leemann & Fabbri *Helv Chim Acta* **4** 2696 1959, Beilstein **25** III/IV 935.]

**Isonicotinamide** [1453-82-3] **M** 122.1, **m** 155.5-156°, **pK<sub>1</sub><sup>20</sup>** -1.0 (protonation of CONH<sub>2</sub>), **pK<sub>2</sub><sup>20</sup>** 3.61, **pK<sub>3</sub><sup>25</sup>** 11.47 (acidic CONH<sub>2</sub>). Recrystallise isonicotinamide from hot water or isopropanol (158.5-159°), and dry it in a vacuum at 100°. The *picrate* crystallises from aqueous EtOH or H<sub>2</sub>O and has **m** 217-218° (214-215°). [Beilstein **22** III/IV 527, **22/2** V 195.]

**Isonicotinic acid (pyridine-4-carboxylic acid)** [55-22-1] M 123.1, m 320°, 323-325°(dec), pK<sub>1</sub><sup>25</sup> 1.70, pK<sub>2</sub><sup>25</sup> 4.89. Crystallise the acid repeatedly from water and dry it under vacuum at 110° or sublime it at 260°/15mm (m 319°). [Beilstein 22 III/IV 518, 22/2 V 188.]

**Isonicotinic acid hydrazide (isoniazide)** [54-85-3] M 137.1, m 172°, pK<sub>1</sub> 1.75 (NHNH<sub>2</sub>), pK<sub>2</sub> 3.57 (=N-), pK<sub>3</sub> 10.75 (-NH). Crystallise isoniazide from 95% EtOH and dry it in a vacuum. [Beilstein 22 III/IV 545, 22/2 V 219.]

**1-Isonicotinyl-2-isopropylhydrazide (Iproniazid)** [54-92-2] M 179.2, m 112.5-113.5°, 114-115°, pK<sub>Est</sub> ~3.5. Crystallise it from \*benzene or \*benzene/petroleum ether and dry it in a vacuum. It is soluble in H<sub>2</sub>O and EtOH. [Fox & Gibas *J Org Chem* 18 994 1953.] The *dihydrochloride* has m 227-228° (from EtOH). [Beilstein 22 III/IV 551.] It is an antidepressant.

**1-Isonicotinyl-2-isopropylhydrazide phosphate (Iproniazid phosphate, Marsilide)** [305-33-9] M 277.2, m 178-179°, 180-184°, pK<sub>Est</sub> ~3.5 (free base). Crystallise it from H<sub>2</sub>O and Me<sub>2</sub>CO. The *free base* (see above) has m 113-114° from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and is soluble in H<sub>2</sub>O and EtOH. [Fox & Gibas *J Org Chem* 18 994 1953, Beilstein 22 III/IV 551.]

**1-Isonicotinyl-2-salicylidenehydrazide** [495-84-1] M 241.2, m 232-233°. Crystallise it from EtOH (m 265-266° or 244-245°), aqueous EtOH or MeOH (252-253°) and dry it in a vacuum at 100°. [Beilstein 22 III/IV 584.]

**5-Isonitrosobarbituric acid (violuric acid)** [26851-19-9] M 175.1, m 221-223°, 245-250°, pK<sub>1</sub> 4.41, pK<sub>2</sub> 9.66 (10.1). Crystallise violuric acid from water or EtOH. *1,1-Dimethylvioluric acid*, m 144-147° has pK<sub>25</sub> 4.72 [Taylor & Robinson *Talanta* 8 518 1961]. [Beilstein 24 III/IV 2142.]

**N-Isopropylcarbazole** [1484-09-9] M 209.3, m 120°. Crystallise it from isopropanol. It sublimes under vacuum. It was also purified by zone refining. The *picrate* has m 143° after recrystallisation from EtOH. [Beilstein 20 I 164.]

**Isoquinoline** [119-65-3] M 129.2, m 24°, 25.5-26°, b 120°/18mm, 243.25°/760mm, d<sub>4</sub><sup>20</sup> 1.0986, n<sub>D</sub><sup>20</sup> 1.6148, pK<sub>25</sub> 5.40. Dry isoquinoline with Linde type 5A molecular sieves or Na<sub>2</sub>SO<sub>4</sub> and fractionally distil at reduced pressure. *Alternatively*, it can be refluxed with, and distilled from, BaO. It is also purified by fractional crystallisation from the melt and distilled from zinc dust. It forms a *phosphate* (m 135°) and a *picrate* (m 223°), which are purified by crystallisation, and the free base can be recovered and distilled. [Packer et al. *J Am Chem Soc* 80 905 1958.] The procedure for purification *via* the picrate comprises the addition of quinoline to picric acid dissolved in the minimum volume of 95% EtOH to yield yellow crystals which are washed with EtOH and air dried before recrystallising from acetonitrile. The crystals are dissolved in dimethyl sulfoxide (previously dried over 4A molecular sieves) and passed through a basic alumina column, on which picric acid is adsorbed. The free base in the effluent is extracted with *n*-pentane and distilled under vacuum. Traces of solvent from small quantities are removed by vapour phase chromatography. The *hydrochloride* crystallises from EtOH with m 193°. [Mooman & Anton *J Phys Chem* 80 2243 1976, Beilstein 20 II 236, 20 III/IV 3410, 20/7 V 333.]

**Isoxanthopterin (2-amino-4,7-dihydroxypteridine)** [529-69-1] M 179.4, m >300°, pK<sub>1</sub><sup>20</sup> -0.5 (basic), pK<sub>2</sub><sup>20</sup> 7.34 (acidic), pK<sub>3</sub><sup>20</sup> 10.06 (acidic). Purify it by repeated precipitation from alkaline solution with acid (preferably AcOH or formic acid), filter, wash well with H<sub>2</sub>O, then EtOH and dry at 100°. The purity is checked by paper chromatography [R<sub>F</sub> 0.15 (*n*-BuOH/AcOH/H<sub>2</sub>O, 4:1:1); 0.33 (3% aqueous NH<sub>4</sub>OH)]. [Goto et al. *Arch Biochem Biophys* 111 8 1965.] [For biochemistry see Blakley *Biochemistry of Folic Acid and Related Pteridines* North Holland Publ Co, Amsterdam 1969.] [Beilstein 26 III/IV 3999.]

**Janus Green B (3-dimethylamino-7-[4-dimethylaminoazo]-5-phenylphenazonium chloride)** [2869-83-2] M 511.1, m >200°, CI 11050. The dye dissolves in H<sub>2</sub>O to give a bluish violet solution which becomes colourless when made 10M in NaOH. It dissolves in EtOH to give a blue-violet colour, filter from insoluble

material, then add dry Et<sub>2</sub>O whereby the dye separates out leaving a small amount of blue colour in solution. Filter off the solid and dry it in a vacuum. Store it in a dark bottle. [Colour Index Vol 4, 3<sup>rd</sup> edn, 4015 1971.]

**Jervine** (3β,23β-17,23-epoxy-2-hydroxyvertraman-11-one, a steroidal alkaloid) [469-59-0] M 425.6, m 243-245°, 247-248°, [α]<sub>D</sub><sup>20</sup> -150° (in EtOH), pK<sub>Est</sub>~9.4. Crystallise Jervine from MeOH/H<sub>2</sub>O or Me<sub>2</sub>O. The *hydrochloride* has m 300-302° (from MeOH/Et<sub>2</sub>O), and the *picrate* has m 262.5° (from aqueous MeOH). [Kutney et al. *Can J Chem* 53 1796 1975, *Beilstein* 27 III/IV 3590.] It is **teratogenic**.

**Julolidine** (2,3,6,7-tetrahydro-1H,5H-benzo[*ij*]quinolizidine) [479-59-4] M 173.3, m 34-36°, 40°, b 105-110°/1mm, 155-156°/17mm, 280°(dec), pK<sub>Est</sub>~7.0. Purify julolidine by dissolving it in dilute HCl, steam is bubbled through the solution and the residual acidic solution is basified with 10N NaOH, extracted with Et<sub>2</sub>O, washed with H<sub>2</sub>O, dried (NaOH pellets), filtered, evaporated and distilled *in vacuo*. The distillate crystallises on cooling (m 39-40°). It develops a red colour on standing in contact with air for several days. The colour can be removed by distilling or dissolving in 2-3 parts of hexane, adding charcoal, filtering and cooling in an Me<sub>2</sub>CO/Dry-ice bath when julolidine crystallises out (85-90% yield m 39-40°). The *hydrobromide* [83646-41-7] has m 218° (239-242°), the *picrate* has m 174°(165°) and the *methiodide* crystallises from MeOH, with m 186° [Glass & Weisberger *Org Synth Coll Vol III* 504 1955, Smith & Yu *J Org Chem* 17 1285 1952, *Beilstein* 20 H 332, 20 I 133, 20 II 214, 20 III/IV 3281.] **Highly TOXIC**.

**Kainic acid H<sub>2</sub>O** (2S,3S,4S-2-carboxy-4-isoprenyl-3-pyrrolidine- acetic acid) [487-79-6] M 231.4, m 235-245°(dec), 251°(dec), [α]<sub>D</sub><sup>20</sup> -14.6° (c 1.46, H<sub>2</sub>O), pK<sub>1</sub> 2.09, pK<sub>2</sub> 4.58, pK<sub>3</sub> 10.21. Purify the acid by adsorbing on to a strongly acidic ion-exchange resin (Merck), elute the diacid with aqueous M NaOH, the eluate is evaporated, H<sub>2</sub>O is added, and filtered through a weakly acidic ion-exchange resin (Merck). The filtrate is then evaporated and recrystallised from EtOH. Its solubility is 0.1g in 1ml of 0.5N HCl. (±)-α-Kainic acid is recrystallised from H<sub>2</sub>O with m 230-260°. Its UV (MeOH) has λ<sub>max</sub> at 219 (log ε 3.9); the <sup>1</sup>HNMR (CCl<sub>4</sub>, 100MHz, Me<sub>4</sub>Si standard) has δ at 1.64 (s 1H), 1.70 (s 3H), 3.24 (d J = 7.5, 2H), 3.3-4.2 (1H), 3.70 (s 3H), 3.83 (s 3H), 4.35 (dd J = 7.5, 14.5, 1H), 5.21 (t J = 7.5, 1H), 7.26 (t J = 7.5, 1H). [Oppolzer & Andres *Helv Chim Acta* 62 2282 1979, *Beilstein* 22 III/IV 1523.]

**Ketanserin** [3(4-*p*-fluorobenzoylpiperidinyl)-*N*-ethylquinazolin-2,4-dione] [74050-98-9] M 395.4, m 227-235°, pK<sup>25</sup> 7.5. Its solubility is 0.001% in H<sub>2</sub>O, 0.038% in EtOH and 2.34% in Me<sub>2</sub>NCHO. It has been purified by recrystallisation from 4-methyl-3-pentanone [Peeters et al. *Cryst Structure Commun* 11 375 1982, Kacprowicz et al. *J Chromatogr* 272 417 1983, Davies et al. *J Chromatogr* 275 232 1983]. It is an antihypertensive.

**Khellin** (4,9-dimethoxy-7-methyl-5-oxofuro[3,2-*g*]-1,2-chromene) [82-02-0] M 260.3, m 154-155°, b 180-200°/0.05mm. Crystallise khellin from H<sub>2</sub>O, MeOH, petroleum ether or Et<sub>2</sub>O. The *hydrochloride* has m 98°(dec) (from EtOH/HCl). [*Beilstein* 19 II 236, 19 III/IV 2816, 19/6 V 320.]

**Kojic acid** [(5-hydroxy-2-hydroxymethyl)-4H-pyran-4-one] [501-30-4] M 142.1, m 152°, 154-155°, pK<sub>1</sub><sup>25</sup> -1.38, pK<sub>2</sub><sup>25</sup> 7.66. Crystallise the acid from MeOH (charcoal) by adding Et<sub>2</sub>O. It sublimes at 150-200°/0.1 torr. [*Beilstein* 18 II 57, 18 III/IV 1145, 18/2 V 516.]

**Kynurenic acid** (4-hydroxyquinoline-2-carboxylic acid) [492-27-3] M 189.1, m 282-283°, 285°, pK<sub>Est(1)</sub>~2, pK<sub>Est(2)</sub>~10. Crystallise the acid from absolute EtOH. The *methyl ester* crystallises from MeOH with m 224-226°. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p2723 1961., *Beilstein* 22 II 174, 22 III/IV 2245, 22/6 V 280.]

**L-Kynurenine** [343-65-7] M 208.2, m 190°(dec), 210°(dec), [α]<sub>D</sub><sup>20</sup> -30° (c 0.4, H<sub>2</sub>O), pK<sub>Est(1)</sub>~2.3, pK<sub>Est(2)</sub>~3.5, pK<sub>Est(3)</sub>~9.2. Crystallise it from H<sub>2</sub>O or aqueous AcOH. The *picrate* has m 188.5-189°(dec) after crystallisation from H<sub>2</sub>O. (±)-Kynurenine has m 218°. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p2726 1961, *Beilstein* 14 III 1657, 14 IV 2562.]

**L-Kynurenine sulfate** [16055-80-4] **M 306.3, m 194°**, monohydrate **m 178°**,  $[\alpha]_{\text{D}}^{25} +9.6^{\circ}$  ( $\text{H}_2\text{O}$ ). Crystallise the sulfate from water by addition of EtOH. The ( $\pm$ )-sulfate has **m 173°**. [Beilstein 14 IV 2562.]

**dl- and l-Laudanosine** {1-[(3,4-dimethoxyphenyl)methyl]-1,2,3,4-tetrahydro-6,7-dimethoxy-2-methylisoquinoline} [(+)-1699-51-0, (-)-2688-77-9] **M 357.4, m 114-115°, 118°**. Crystallise these from EtOH. The ( $\pm$ )-picrate crystallises from EtOH with **m 177-178°**. The (-)-isomer has **m 83-85°** and  $[\alpha]_{\text{D}}^{24} -110^{\circ}$  (c 0.3, EtOH). The hydrobromide tetrahydrate [303136-74-5] **M 436.3** has **m 232-234°**. [Frydman et al. *Tetrahedron* 4 342 1958, Elliott *J Heterocycl Chem* 9 853 1972, *Beilstein* 21 II 183, 21 III/IV 2704.]

**Lumichrome (7,8-dimethylalloxazine)** [1086-80-2] **M 242.2, m >290°, pK<sub>Est(1)</sub> ~0.1 (basic), pK<sub>Est(2)</sub> ~9.9 (acidic)**. Recrystallise lumichrome twice from glacial AcOH and dry it at 100° in a vacuum. [Cresswell & Wood *J Chem Soc* 4768 1960, *Beilstein* 26 III/IV 2538.]

**Luminol (5-aminophthalazin-1,4-dione)** [521-31-3] **M 177.2, m 329-332°, pK<sub>1</sub> 3.37, pK<sub>2</sub> 6.35**. Dissolve luminol in KOH solution, treat with Norit (charcoal), filter and precipitate it with conc HCl. [Hardy et al. *Talanta* 24 297 1977.] Store it in the dark in an inert atmosphere, because its structure changes during its luminescence. It has been recrystallised from 0.1M KOH [Merenyi et al. *J Am Chem Soc* 108 77716 1986]. [Beilstein 25 II 389, 25 III/IV 4192.]

**dl-Lupinine (1-hydroxymethyloctahydroquinolizine, 1-hydroxymethylquinolizidine)** [10248-30-3] **M 169.3, m 57-58°, 59°, 107°/1mm, pK<sub>Est</sub> ~7.0**. It crystallises from Me<sub>2</sub>CO, pentane and petroleum ether (b. 40-60°) and can be sublimed or distilled in high vacuum. The picrate has **m 127°**(from Et<sub>2</sub>O), the picrolonate has **m 203-204°**, and the methiodide has **m 203°**(dec, from EtOH). [Clemo et al. *J Chem Soc* 969 1937, Boekelheide & Lodge *J Am Chem Soc* 73 3684 1951, *Beilstein* 21 II 28, 21 III/IV 291.]

**Lutidine (mixture)**. For the preparation of pure 2,3-, 2,4- and 2,5-lutidine from commercial “2,4- and 2,5-lutidine” see Coulson et al. *J Chem Soc* 1934 1959, and Kyte, Jeffery and Vogel *J Chem Soc* 4454 1960.

**2,3-Lutidine** [583-61-9] **M 107.2, f -14.8°, b 160.6°, d<sub>4</sub><sup>20</sup> 0.9464, n<sub>D</sub><sup>20</sup> 1.50857, pK<sup>25</sup> 6.57**. Steam distil it from a solution containing about 1.2 equivalents of 20% H<sub>2</sub>SO<sub>4</sub>, until ca 10% of the base has been carried over with the non-basic impurities. The acidic solution is then made alkaline, and the base is separated, dried over NaOH or BaO, and fractionally distilled. The distilled lutidine is converted to its urea complex by stirring 100g with 40g of urea in 75ml of H<sub>2</sub>O, cooling to 5°, filtering at the pump, and washing with 75ml of H<sub>2</sub>O. The complex, dissolved in 300ml of H<sub>2</sub>O, is steam distilled until the distillate gives no turbidity with a little solid NaOH. The distillate is then treated with excess solid NaOH, and the upper layer is removed: the aqueous layer is then extracted with diethyl ether. The upper layer and the ether extract are combined, dried (K<sub>2</sub>CO<sub>3</sub>), and distilled through a short column. Final purification is by fractional crystallisation using partial freezing. The picrate crystallises from EtOH with **m 187-188°**. [Kyte et al. *J Chem Soc* 4454 1960, *Beilstein* 20 H 243, 20 II 159, 20 III/IV 2765, 20/6 V 15.]

**2,4-Lutidine** [108-47-4] **M 107.2, b 157.8°, d<sub>4</sub><sup>20</sup> 0.9305, n<sub>D</sub><sup>20</sup> 1.50087, n<sub>D</sub><sup>25</sup> 1.4985, pK<sup>25</sup> 6.77**. Dry it with Linde type 5A molecular sieves, BaO (*do not use Na, see picolines below*), and fractionally distil it. The distillate (200g) is heated with \*benzene (500ml) and conc HCl (150ml) in a Dean and Stark apparatus on a water bath until water no longer separates, and the temperature just below the liquid reaches 80°. When cold, the supernatant \*benzene is decanted, and the 2,4-lutidine hydrochloride, after washing with a little \*benzene, is dissolved in water (350ml). After removing any \*benzene by steam distillation, an aqueous solution of NaOH (80g) is added, and the free lutidine is steam distilled. It is isolated by saturating the distillate with solid NaOH and distilling it through a short column. The precipitation cycle is repeated, then the final distillate is partly frozen in an apparatus at -67.8-68.5° (cooled by acetone/CO<sub>2</sub>). The crystals are collected, then melted and distilled. [Kyte et al. *J Chem Soc* 4454 1960.] *Alternative purifications are via the picrate m 183-184°* (from H<sub>2</sub>O). [Clarke & Rothwell *J Chem Soc* 1885 1960], or the hydrobromide [Warnhoff *J Org Chem* 27 4587 1962]. The latter is precipitated from a solution of lutidine in \*benzene by passing dry HBr gas: the salt is recrystallised from CHCl<sub>3</sub>/methyl ethyl ketone, then decomposed with NaOH, and the free base is extracted

into Et<sub>2</sub>O, dried, evaporated and the residue is distilled. [*Beilstein* 20 II 180, 20 III/IV 2718, 20/6 V 19.]

**2,5-Lutidine** [589-93-5] M 107.2, m -15.3°, b 156.7°/759mm, d<sub>4</sub><sup>20</sup> 0.927, n<sub>D</sub><sup>25</sup> 1.4982, pK<sup>25</sup> 6.40. Steam distil the lutidine from a solution containing 1-2 equivalents of 20% H<sub>2</sub>SO<sub>4</sub> until about 10% of the base has been carried over with the non-basic impurities, then the acidic solution is made alkaline, and the base is separated, dried with NaOH and fractionally distilled twice. Dry the distillate with Na and fractionally distil it through a Todd column packed with glass helices. The *hydrochloride* has m 219°(dec), and the *picrate* has m 170.5° (from EtOH or H<sub>2</sub>O). [*Beilstein* 20 H 244, 20 II 160, 20 III/IV 2774, 20/6 V 27.]

**2,6-Lutidine** [108-48-5] M 107.2, m -59°, b 144.0°, d<sub>4</sub><sup>20</sup> 0.92257, n<sub>D</sub><sup>20</sup> 1.49779, pK<sup>25</sup> 6.72. Likely contaminants include 3- and 4-picoline (similar boiling points). However, they are removed by using BF<sub>3</sub>, with which they react preferentially, by adding 4ml of BF<sub>3</sub> to 100ml of dry fractionally distilled 2,6-lutidine and redistilling. Distillation of commercial material from AlCl<sub>3</sub> (14g per 100ml) can also be used to remove picolines (and water). *Alternatively*, lutidine (100m) can be refluxed with ethyl benzenesulfonate (20g) or ethyl *p*-toluenesulfonate (20g) for 1 hour, then the upper layer is cooled, separated and distilled. The distillate is refluxed with BaO or CaH<sub>2</sub>, then fractionally distilled through a glass helices-packed column.

2,6-Lutidine can be dried with KOH (do not use Na, see picolines below) or by refluxing with (and distilling from) BaO, prior to distillation. For purification *via* its *picrate*, 2,6-lutidine, dissolved in absolute EtOH, is treated with an excess of warm ethanolic picric acid. The precipitate is filtered off, recrystallised from acetone (to give m 163-164.5° (166-167°), and partitioned between ammonia and CHCl<sub>3</sub>/diethyl ether. The organic layer, after washing with dilute aqueous KOH, is dried with Na<sub>2</sub>SO<sub>4</sub> and fractionally distilled. [Warnhoff *J Org Chem* 27 4587 1962.] *Alternatively*, 2,6-lutidine can be purified *via* its urea complex, as described under 2,3-lutidine. Other purification procedures include azeotropic distillation with phenol [Coulson et al. *J Appl Chem (London)* 2 71 1952], fractional crystallisation by partial freezing, and vapour-phase chromatography using a 180-cm column of polyethylene glycol-400 (Shell, 5%) on Embacel (May and Baker) at 100°, with argon as carrier gas [Bamford & Block *J Chem Soc* 4989 1961]. The *hydrochloride* has m 235-237°, 239° (from EtOH). [*Beilstein* 20 II 160, 20 III/IV 2776, 20/6 V 32.]

**3,5-Lutidine** [591-22-0] M 107.2, f -6.3°, b 172.0°/767mm, d<sub>4</sub><sup>20</sup> 0.9419, n<sub>D</sub><sup>20</sup> 1.50613, n<sub>D</sub><sup>25</sup> 1.5035, pK<sup>25</sup> 6.15. Dry 3,5-lutidine with CaH<sub>2</sub> and fractionally distil it through a Todd column packed with glass helices. Dissolve (100ml) in dilute HCl (1:4) and steam distil this until 1L of distillate is collected. Excess conc NaOH is added to the residue which is again steam distilled. The base is extracted from the distillate, using diethyl ether. The extract is dried over K<sub>2</sub>CO<sub>3</sub>, and distilled. It is then fractionally crystallised by partial freezing. The *hydrochloride* has m 229°(sublimes at 190-231°), and the *picrate* has m 242-243°(dec, from H<sub>2</sub>O), 249-250°(dec, from AcOH). [*Beilstein* 20 II 161, 20 III/IV 2788, 20/6 V 60.]

**Lycorine** [476-28-8] M 552.9, m 275-280°(dec), b 175-185°/0.002mm. [α]<sub>D</sub><sup>20</sup> -91° (c 0.16, EtOH), pK<sup>25</sup> 6.9 (30% aqueous dimethylformamide). It crystallises as orange crystals from MeOH (m 281-283°), CHCl<sub>3</sub>/EtOH (m 272-274°), pyridine or from EtOH (m 277° dec). It has been distilled under high vacuum. The *hydrochloride* has m 288° (from MeOH/HCl), and the *picrate* has m 196-197°(from EtOH), [Cook et al. *J Chem Soc* 4176 1954, Martin & Tu *J Org Chem* 46 3763 1981, *Beilstein* 27 II 547, 27 III/IV 6463.]

(+)-Lysergic acid [82-58-6] M 268.3, m 240°(dec), [α]<sub>D</sub><sup>20</sup> +40° (pyridine), pK<sub>1</sub><sup>25</sup> 3.32, pK<sub>2</sub><sup>25</sup> 8.66, pK<sup>20</sup> 8.50, pK<sup>40</sup> 8.27. It crystallises from water as a *hydrate*. The *methyl ester* crystallises from \*C<sub>6</sub>H<sub>6</sub> and has m 168°; the *amide* [478-94-4] has m 242°(dec) (from MeOH) and [α]<sub>546</sub> +15° (c 0.5, pyridine). The (-)-*hydrochloride* has m 208-210°(dec, from MeOH). [Kornfeld et al. *J Am Chem Soc* 76 5256 1954, Kornfeld et al. *J Am Chem Soc* 78 3087 1956, *Beilstein* 25 III/IV 934.]

**Maltol (3-hydroxy-2-methyl-4-pyrone)** [118-71-8] M 126.1, m 161-162°, 162-162.5°. It crystallises from CHCl<sub>3</sub>, toluene, aqueous 50% EtOH or H<sub>2</sub>O, and is volatile in steam. It can be readily sublimed in a vacuum. It forms a Cu<sup>2+</sup> complex. [*Beilstein* 17 III/IV 5916, 18/1 V 114.]

**Meconic acid (3-hydroxy-γ-pyrone-2,6-dicarboxylic acid)** [497-59-6] M 200.1, m 100° (loses H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup>

**1.83, pK<sub>2</sub><sup>25</sup> 2.3, pK<sub>3</sub><sup>20</sup> 10.10.** Crystallise the acid from water (0.25g/ml) and dry it at 100° for 20 minutes to dehydrate the *mono* or *dihydrate*. It decarboxylates above 120° or in boiling H<sub>2</sub>O. It is soluble in MeOH (2%), EtOAc (2%) and Me<sub>2</sub>CO (1%). The *picrate* has **m** 206.5-208.5°(dec, from H<sub>2</sub>O). [Wibaut & Kleinpool *Rec Trav Chim Pays Bas* **66** 24 1947, *Beilstein* **18** H 409, **18** I 523, **18** II 367, **18** III/IV 6136.]

**Melamine (2,4,6-triamino-1,3,5-triazine)** [108-78-1] **M 126.1, m 353°**, **pK<sup>25</sup> 5.00.** Crystallise Melamine from water or dilute aqueous NaOH. It sublimes at ~240° on prolonged heating. [*Beilstein* **26** I 74, **26** II 132, **26** III/IV 1253.]

**(±)-Mellein [(±)-3,4-dihydro-8-hydroxy-3-methyl-2-benzopyran-1-one, 8-hydroxy-3-methylisochroman-1-one]** [1200-93-7] **M 178.2, m 37-39°, 39°, pK<sub>Est</sub> ~9.5.** Purify it by recrystallisation from H<sub>2</sub>O or aqueous EtOH. It has UV with λ<sub>max</sub> at 247 and 314nm. [Arakawa et al. *Justus Liebigs Ann Chem* **728** 152 1969, Blair & Newbold *Chem Ind (London)* 93 1955, *J Chem Soc* 2871 1955.] The *methyl ether* has **m** 66-67° and UV with λ<sub>max</sub> at 242nm (ε 7,400) and 305nm (ε 4,600). *R(-)-Mellein* has **m** 56°(aqueous Me<sub>2</sub>CO), [α]<sub>D</sub><sup>25</sup> -102.5° (c 1, CHCl<sub>3</sub>) and *R(+)-mellein* has **m** 56-57°(hexane), [α]<sub>D</sub><sup>25</sup> +102° (c 1, CHCl<sub>3</sub>) or [α]<sub>D</sub><sup>25</sup> +88° (c 1, MeOH). [*Beilstein* **18** III/IV 188, **18**/I V 274.]

**2-Mercaptobenzimidazole** [583-39-1] **M 150.2, m 302-304°, 312°, pK<sup>20</sup> 10.24.** Crystallise it from aqueous EtOH, AcOH or aqueous ammonia. It complexes with many metals. [Brown *J Chem Soc* 1976 1958, *Beilstein* **24** II 65, **24** III/IV 287.]

**2-Mercaptobenzothiazole** [149-30-4] **M 167.2, m 182°, pK<sup>25</sup> 7.5 (50% aqueous AcOH).** Crystallise it repeatedly from 95% EtOH, or purify it by incomplete precipitation by dilute H<sub>2</sub>SO<sub>4</sub> from a basic solution, followed by several crystallisations from acetone/H<sub>2</sub>O or \*benzene. It complexes with Ag, Au, Bi, Cd, Hg, Ir, Pt, and Tl. [*Beilstein* **27** II 233, **27** III/IV 2709.]

**2-Mercaptoimidazole** [872-35-5] **M 100.1, m 221-222°, 226-228°(monohydrate), pK<sub>1</sub><sup>20</sup> -1.6, pK<sub>2</sub><sup>20</sup> 11.6.** Crystallise 2-mercaptoimidazole from Me<sub>2</sub>CO or H<sub>2</sub>O. Its UV has λ<sub>max</sub> at 208 and 252nm (H<sub>2</sub>O). [Fox et al. *J Am Chem Soc* **67** 496 1947, *Beilstein* **24** II 7, **24** III/IV 61.]

**2-Mercapto-1-methylimidazole** [60-56-0] **M 114.2, m 145-147°, 146-148°, pK<sub>1</sub><sup>20</sup> -2.0, pK<sub>2</sub><sup>20</sup> 11.9.** Crystallise it from EtOH. Its UV has λ<sub>max</sub> at 251nm (H<sub>2</sub>O), 260nm (EtOH) and 267nm (CHCl<sub>3</sub>). [Lawson & Morley *J Chem Soc* 1103 1956, *Beilstein* **24** H 17, **24** III/IV 61.]

**6-Mercaptopurine monohydrate** [6112-76-1] **M 170.2, m 314-315°(dec), ~315°(dec), 313-315°(dec), pK<sub>1</sub><sup>20</sup> 0.5, pK<sub>2</sub><sup>20</sup> 7.77, pK<sub>3</sub><sup>20</sup> 10.84.** Crystallise 6-mercaptoquinoline from pyridine (30ml/g), wash it with pyridine, then triturate with water (25ml/g) and adjust to pH 5 by adding M HCl. Recrystallise it by heating, then cooling, the solution. Filter off the solid, wash it with water and dry it at 110°. It has also been crystallised from water (charcoal) as yellow crystals of the *monohydrate* which become *anhydrous* on drying at 140°. It has UV with λ<sub>max</sub> at 230 and 312nm (ε 14,000 and 19,600) in 0.1N NaOH; 222 and 327nm (ε 9,2400 and 21,300), and 216 and 329nm (ε 8,740 and 19,300) in MeOH. It forms a 1:1 complex with Zn<sup>2+</sup>, Pb<sup>2+</sup>, Co<sup>2+</sup>, and Ni<sup>2+</sup> in aqueous dioxan. It is an antineoplastic. [Albert & Brown *J Chem Soc* 2060 1954, IR: Brown & Mason *J Chem Soc* 682 1957, UV: Fox et al. *J Am Chem Soc* **80** 1669 1958, UV: Mason *J Chem Soc* 2071 1954, *Beilstein* **26** III/IV 2097.]

**8-Mercaptoquinoline (2H<sub>2</sub>O, thioxine)** [491-33-8] **M 197.3, m 58-59°, pK<sub>1</sub><sup>25</sup> 2.0, pK<sub>2</sub><sup>25</sup> 8.40.** Thioxine readily oxidises in air to give diquinolyl-8,8'-disulfide (which is stable). It is more convenient to make 8-mercaptoquinoline by reduction of the disulfide. [Nakamura & Sekido *Talanta* **17** 515 1970.] The *hydrochloride* (see thioxine hydrochloride below) is more stable. [*Beilstein* **21** III/IV 1197, **21**/3 V 30.]

**3-Methoxycarbonyl-2,5-dihydrothiophen-1,1-dioxide (methyl 3-sulfolene-3-carboxylate)** [67488-50-0] **M 176.1, m 57-58°, 60-62°.** If the IR shows OH bands, then dissolve the dioxide in CHCl<sub>2</sub>, wash it with aqueous Na<sub>2</sub>CO<sub>3</sub> and H<sub>2</sub>O, dry it over MgSO<sub>4</sub>, filter, evaporate and wash the residue with cold Et<sub>2</sub>O and dry *in vacuo*. Its <sup>1</sup>HNMR (CDCl<sub>3</sub>) has δ at 7.00 (m 1H), 3.98 (bs 4H) and 3.80 (s Me). [McIntoch & Sieber *J Org Chem* **43**

4431 1978, *Beilstein* **18/6** V 5.]

**5-Methoxyindole** [1006-94-6] **M 147.2**, **m 55°, 57°**, **b 176-178°/17mm**, **pK<sub>Est</sub> ~0**. Crystallise 5-methoxyindole from cyclohexane petroleum ether or petroleum ether/Et<sub>2</sub>O. [Saito & Kikugawa *J Heterocycl Chem* **16** 1325 1979, *Beilstein* **21** III/IV 765, **21/3** V 18.]

**5-(*p*-Methoxyphenyl)-1,2-dithiole-3-thione** [42766-10-9] **M 240.2**, **m 111°**. Crystallise the thione from EtOAc, BuOAc or EtOH. It sublimes at 90°/0.001mm and complexes with Sb<sup>3+</sup>, Sb<sup>5+</sup>, Bi<sup>3+</sup>, Sn<sup>4+</sup>, Ag<sup>+</sup>, Au<sup>3+</sup> and Hg<sup>2+</sup>. [*Beilstein* **19** III/IV 2538.]

**8-Methoxypsoralen (Methoxsalen)** See xanthotoxin in “Miscellaneous Compounds” in Chapter 7.

**6-Methylaminopurine** [443-72-1] **M 149.2**, **m >300°, 312-314° (dec)**, **pK<sub>1</sub><sup>20</sup> <1**, **pK<sub>2</sub><sup>20</sup> 4.15**, **pK<sub>3</sub><sup>20</sup> 10.02**. The purine is best purified by recrystallising 2g from 50ml of H<sub>2</sub>O and 1.2g of charcoal. [UV: Albert & Brown *J Chem Soc* 2060 1954; UV: Mason *J Chem Soc* 2071 1954; see also Elion et al. *J Am Chem Soc* **74** 411 1952.] The *picrate* has **m 265°(257°)** [Bredereck et al. *Chem Ber* **81** 307 1948]. [*Beilstein* **26** III/IV 3565.]

**Methyl 3-aminopyrazine-2-carboxylate** [16298-03-6] **M 153.1**, **m 169-172°, 172°**. The ester forms yellow needles from H<sub>2</sub>O (100 parts using charcoal). If it contains the free acid (see IR), then dissolve it in CH<sub>2</sub>Cl<sub>2</sub>, wash it with saturated aqueous Na<sub>2</sub>CO<sub>3</sub>, brine, dry over MgSO<sub>4</sub> filter, evaporate and recrystallise the residue. The *free acid* has **m 203-204° (dec)** [UV: Brown & Mason *J Chem Soc* 3443 1956] with **pK<sub>1</sub> <1** and **pK<sub>2</sub> 3.70**. The *ammonium salt* has **m 232°(dec)** (from aqueous Me<sub>2</sub>CO) and the *amide* has **m 239.2°** (from H<sub>2</sub>O) [Ellingson et al. *J Am Chem Soc* **67** 1711 1945]. [*Beilstein* **25** III/IV 4412.]

**9-Methylcarbazole (*N*-methylcarbazole)** [1484-12-4] **M 181.2**, **m 87°, 89°**. Purify *N*-methylcarbazole by chromatography on silica gel and eluting with CHCl<sub>3</sub>/Me<sub>2</sub>CO/Et<sub>2</sub>O (100:5:1 v/v), or by flash chromatography using petroleum ether, or by zone melting followed by recrystallisation from petroleum ether or EtOH. [Flo & Pindus *Annalen* 509 1987, Kashima et al. *J Heterocycl Chem* **24** 913 1987, UV: Armarego *Physical Methods in Heterocyclic Chemistry* (Ed Katritzky, Academic Press) **Vol III** 158 1971, *Beilstein* **20** H 436.]

**5-Methylcytosine [4-amino-5-methylpyrimidin-2(1*H*)-one]** [554-01-8] **M 125.1**, **m 270°(dec)**, **pK<sub>1</sub> 4.6**, **pK<sub>2</sub> 12.4**. Crystallise it from water (solubility is 3.4%). The *hydrochloride* has **m 299-301°** (sintering at 280°) (from aqueous HCl/Me<sub>2</sub>CO). [Hitchings et al. *J Biol Chem* **177** 537 1949, Cohn *J Am Chem Soc* **73** 1539 1951, *Beilstein* **25** II 183, **25** III/IV 3727.]

**2-Methyl-1,3-dithiane** [6007-26-7] **M 134.3**, **b 53-54°/1.1mm**, **66°/5mm**, **79-80°/8-10mm**, **85°/12mm**, **d<sub>4</sub><sup>20</sup> 1.121**, **n<sub>D</sub><sup>20</sup> 1.560**. Wash the dithiane with H<sub>2</sub>O, 2.5 M aqueous NaOH, H<sub>2</sub>O, brine, dry over K<sub>2</sub>CO<sub>3</sub> (use toluene as solvent if the volume of reagent is small), filter, evaporate and distil the colourless residue. Its IR film has  $\nu_{\max}$  at 1455, 1371 and 1060 (all medium and CH<sub>3</sub>), 1451m, 1422s, 1412m, 1275m, 1236m, 1190m, 1171w, 918m and 866w (all dithiane) cm<sup>-1</sup> [Corey & Erickson *J Org Chem* **36** 3553 1971, Seebach & Corey *J Org Chem* **40** 231 1975]. [*Beilstein* **19** III/IV 49, **19/1** V 53.]

**Methylene Blue [3,7-bis-(dimethylamino)phenothiazin-5-ium chloride]** [61-73-4] **M 319.9**, **CI 52015**,  $\epsilon_{654}$  **94,000 (EtOH)**,  $\epsilon_{664}$  **81,000 (H<sub>2</sub>O)**, **pK<sup>25</sup> 3.8**. Crystallise the chloride from 0.1M HCl (16ml/g), the crystals are separated by centrifugation, washed with chilled EtOH and diethyl ether, and dried under vacuum. Crystallise it from 50% aqueous EtOH, wash it with absolute EtOH, and dry it at 50-55° for 24 hours. It has also been crystallised from \*benzene/MeOH (3:1). It has been salted out with NaCl from a commercial concentrated aqueous solution, then crystallised from water, and dried at 100° in an oven for 8-10 hours. [*Beilstein* **27** III/IV 5152.]

**Methylene Green [3,7-bis-(dimethylamino)-4-nitrophenothiazin-5-ium chloride]** [2679-01-8] **M 364.9**, **m >200°(dec)**, **CI 52020**, **pK<sup>25</sup> 3.2**. Crystallise the dye three times from water (18ml/g). [*Beilstein* **27** H 399.] The *ZnCl<sub>2</sub> double salt* [6722-15-2] **M 866.0** has the same **CI**. [*Beilstein* **27** III/IV 5157.]

**2-Methylene-oxetan-2-one (diketene)** [674-82-8] M 84.1, m  $-7^{\circ}$ , b  $41^{\circ}/50.5\text{mm}$ ,  $66-68^{\circ}/90\text{mm}$ ,  $127^{\circ}/760\text{mm}$ ,  $d_4^{20}$  1.440,  $n_D^{20}$  1.4376,  $n_D^{25}$  1.4348. Diketene polymerises violently in the presence of alkali. Distil it under reduced pressure, then fractionally crystallise it by partial freezing (using as a cooling bath made from a 1:1 solution of  $\text{Na}_2\text{S}_2\text{O}_3$  in water, cool with Dry-ice until slushy, and store it in a Dewar flask). Freezing proceeds slowly, and takes about a day for half completion. The crystals are separated and stored in a refrigerator under  $\text{N}_2$ . [Miller & Carlson *J Am Chem Soc* 79 3995 1957, Andreades & Carlson *Org Synth Coll Vol V* 679 1973, *Beilstein* 7 H 552, 7 I 309, 7 II 525, 1 III 2947, 17 III/IV 4297, 17/9 V 115.] See ketene [463-51-4] in “Aliphatic Compounds”, in this Chapter.

**2-Methylfuran (Silvan)** [534-22-5] M 82.1, f  $-90.19^{\circ}$ , b  $62.7-62.8^{\circ}/731\text{mm}$ ,  $d_4^{20}$  0.917,  $n_D^{20}$  1.436. Wash it with acidified saturated ferrous sulfate solution (to remove peroxides), separate, dry with  $\text{CaSO}_4$  or  $\text{CaCl}_2$ , and fractionally distil it from KOH immediately before use. To reduce the possibility of spontaneous polymerisation, addition of about one-third of its volume of heavy mineral oil to 2-methylfuran prior to distillation has been recommended. [*Beilstein* 17 H 36, 17 I 18, 17 II 39, 17 III/IV 265.]

**1-Methylguanidine** [938-85-2] M 165.2, m  $>300^{\circ}(\text{dec})$ ,  $\text{pK}_1^{20} \sim 0$ ,  $\text{pK}_2^{20}$  3.13,  $\text{pK}_3^{20}$  10.54. Crystallise it from  $\text{H}_2\text{O}$  or 50% aqueous acetic acid. [*Beilstein* 26 III/IV 3892.]

**7-Methylguanidine** [578-76-7] M 165.2,  $\text{pK}_1^{20} \sim 0$ ,  $\text{pK}_2^{20}$  3.50,  $\text{pK}_3^{20}$  9.95. Crystallise it from water. It has UV with  $\lambda_{\text{max}}$  at 280nm (pH 2.1). The *picrate* has m  $267^{\circ}$  ( $270-272^{\circ}$  dec, also reported). [*Beilstein* 26 H 455, 26 I 134, 26 II 263, 26 III/IV 3890.]

**N-Methylimidazole** [616-47-7] M 82.1, b  $81-84^{\circ}/27\text{mm}$ ,  $197-198^{\circ}/760\text{mm}$ ,  $d_4^{20}$  1.032,  $n_D^{20}$  1.496,  $\text{pK}^{25}$  7.25. Dry it with sodium metal and then distil it. Store it at  $0^{\circ}$  under dry argon. The *picrate* has m  $159.5-160.5^{\circ}$  (from  $\text{H}_2\text{O}$ ). [*Beilstein* 23 III/IV 568.]

**2-Methylimidazole (2-methylglyoxaline)** [693-98-1] M 82.1, m  $140-141^{\circ}$ ,  $144.5-145.5^{\circ}$ , b  $267^{\circ}/760\text{mm}$ ,  $\text{pK}^{25}$  7.86. Recrystallise 2-methylimidazole from \*benzene or petroleum ether. The *picrate* has m  $215^{\circ}$  (from  $\text{H}_2\text{O}$ ). [*Beilstein* 23 III/IV 594, 23/5 V 35.]

**4-Methylimidazole** [822-36-6] M 82.1, m  $47-48^{\circ}$ , b  $263^{\circ}/760\text{mm}$ ,  $\text{pK}^{25}$  7.61. Recrystallise 4-methylimidazole from \*benzene or petroleum ether. It has m  $56^{\circ}$  after sublimation. The *picrate* has m  $162-163.5^{\circ}$  (from EtOH). [*Beilstein* 23 II 60, 23 III/IV 597, 23/5 V 89.]

**2-Methylindole** [95-20-5] M 131.2, m  $61^{\circ}$ ,  $\text{pK}^{25}$   $-0.28$  (C-3 protonation, aqueous  $\text{H}_2\text{SO}_4$ ). Crystallise it from \*benzene. It has also been purified by zone melting. The *picrate* has m  $139^{\circ}$  (from  $\text{Et}_2\text{O}$  or  $\text{Et}_2\text{O}/\text{MeOH}$ ). [Cohen et al. *J Am Chem Soc* 82 2184 1960, *Beilstein* 20 III/IV 3202, 20/7 V 59.]

**3-Methylindole (skatole)** [83-34-1] M 131.2, m  $95^{\circ}$ ,  $\text{pK}^{25}$   $-4.55$  (C-3-protonation, aqueous  $\text{H}_2\text{SO}_4$ ). Crystallise skatole from \*benzene or petroleum ether (m  $96.5^{\circ}$ ). It has also been purified by zone melting. The *picrate* has m  $182^{\circ}$  (from  $\text{Et}_2\text{O}$  or  $\text{Et}_2\text{O}/\text{MeOH}$ ). [*Beilstein* 20 III/IV 3206, 20/7 V 69.]

**N-Methylmorpholine (4-methylmorpholine)** [109-02-4] M 101.2, b  $116-117^{\circ}/764\text{mm}$ ,  $d_4^{20}$  0.919,  $n_D^{20}$  1.436,  $\text{pK}^{25}$  7.41. Dry it by refluxing with BaO or sodium, then fractionally distil it through a helices-packed column. The *picrate* has m  $227^{\circ}$ , the *thiocyanate salt* has m  $103^{\circ}$  (from butanone). [Hall *J Phys Chem* 60 63 1956, *Beilstein* 27 I 203, 27 III/IV 22.]

**4-Methylmorpholine-4-oxide monohydrate** [7529-22-8] M 135.2, m  $71-73^{\circ}$ . When the oxide is dried for 2-3 hours at high vacuum, it dehydrates. Add MeOH to the oxide and distil off the solvent under vacuum until the temperature is *ca*  $95^{\circ}$ . Then add  $\text{Me}_2\text{CO}$  at reflux and cool to  $20^{\circ}$ . The crystals are filtered off, washed with  $\text{Me}_2\text{CO}$  and dried. The degree of hydration may vary and may be important for the desired reactions. [van Rheen et al. *Tetrahedron Lett* 1973 1076, Schneider & Hanze *US Pat* 2 769 823; see also Sharpless et al. *Tetrahedron Lett* 2503 1976.]

**3-Methyl-2-oxazolidone** [19836-78-3] M 101.1, m 15°, b 88-91°/1mm,  $d_4^{20}$  1.172,  $n_D^{20}$  1.455. Purify the oxazolidone by successive fractional freezing, then dry it in a dry-box over 4A molecular sieves for 2 days. Distil it under high vacuum and store it dry as before. [Beilstein 27 III/IV 2517.]

**3-Methyl-3-oxetanemethanol (3-hydroxymethyl-3-methyloxetane)** [3143-02-0] M 102.1, b 80°/4mm, 92-93°/12mm,  $d_4^{20}$  1.033,  $n_D^{25}$  1.4449. Purify the oxetane by fractionation through a glass column [Pattison *J Am Chem Soc* 79 3455 1957, Corey et al. *J Am Chem Soc* 106 2736 1984]. [Beilstein 17 III/IV 1128.]

**5-Methyl-1,10-phenanthroline** [3002-78-6] M 194.2, m 67°(monohydrate), 113°(anhydrous),  $pK^{25}$  5.28. Crystallise it from \*benzene/petroleum ether or from H<sub>2</sub>O as the monohydrate. It complexes with many metals. [Beilstein 23 III/IV 1714.]

**5-Methylphenazinium methyl sulfate** [299-11-6] M 306.3, m 155-157° (198°dec by rapid heating),  $pK^{25}$  -3.5. It forms yellow-brown prisms from EtOH (charcoal), or EtOH/Et<sub>2</sub>O. Its solubility in H<sub>2</sub>O at 20° is 10%. In the presence of aqueous KI it forms a *semiquinone* which crystallises as blue leaflets from EtOH. [Wieland & Roseen *Chem Ber* 48 1117 1913, Voriskova *Col Czech Chem Commun* 12 607 1947, Bülow *Chem Ber* 57 1431 1924, Campbell et al. *J Chem Soc* 404 1938, Morley *J Chem Soc* 4008 1952, Beilstein 23 I 59, 23 II 234, 23 III/IV 1658, 23/8 V 395.]

**N-Methylphenothiazine (10-methylphenothiazine)** [1207-72-3] M 213.2,  $\alpha$ -form m 99.3°, 100-102°, and b 360-365°,  $\beta$ -form m 78-79°. Recrystallise it (three times) from EtOH to give the  $\alpha$ -form (prisms). Recrystallisation from EtOH/\*benzene gives the  $\beta$ -form (needles). It has also been purified by vacuum sublimation and is carefully dried in a vacuum line. It has been crystallised from toluene or MeOH and stored in the dark [Guarr et al. *J Am Chem Soc* 107 5104 1985, Olmsted et al. *J Am Chem Soc* 109 3297 1987]. Its solubility in H<sub>2</sub>O (pH 6.5) is 0.0544mg/100ml, and in hexane it is 2083mg/100ml. Its UV has  $\lambda_{max}$  at 255nm (log $\epsilon$  4.60) (pH 6.3) and 255nm in hexane. [Cymerman-Craig & Warburton *Aust J Chem* 9 294 1956, Beilstein 27 II 33, 27 III/IV 1215.]

**3-Methyl-1-phenyl-5-pyrazolone** [89-25-8] M 174.2, m 127°, 129°,  $pK^{25}$  2.7. Crystallise the pyrazolone from hot H<sub>2</sub>O, EtOH or EtOH/water (1:1). It complexes with metals. [Veibel et al. *Acta Chim Scand* 6 1066 1952, Beilstein 24 II 9, 24 III/IV 71.]

**1-Methyl-4-phenyl-1,2,3,6-tetrahydropyridine hydrochloride (MPTP)** [23007-85-4] M 209.7, m 196-198°,  $pK_{Est}$  ~9.3,  $pK_a$  8.07 (50% aqueous EtOH). Purify MPTP by recrystallisation from Me<sub>2</sub>CO/isoPrOH. The *free base* has m 40-42° (from heptane), b 99-100°/1.3mm, 128-132°/12mm, (137-142°/0.8mm),  $n_D^{25}$  1.5347. The *hydrochloride* has m 251-252° (from Me<sub>2</sub>CO/isoPrOH) [Schmidle & Mansfield *J Am Chem Soc* 78 425 1956, Defeudis *Drug Dev Res* 15 1 1988, Beilstein 20 III/IV 3240, 20/7 V 121.]

**(±)-2-Methylpiperazine** [109-07-9] M 100.2, m 61-62°, 66°, b 147-150°/739mm,  $pK_1^{25}$  5.46,  $pK_2^{25}$  9.90. Purify it by zone melting and by distillation. It is *hygroscopic*. The *picrate* has m 275-276°. [Beilstein 23 II 16, 23 III/IV 393, 23/3 V 267.]

**(±)-3-Methylpiperidine** [626-56-2] M 99.2, b 125°/763mm,  $d_4^{20}$  0.846,  $n_D^{25}$  1.4448,  $pK^{25}$  10.92. Purify it *via* the *hydrochloride* (m 172°). The *hydrobromide* has m 162-163° (from iso-PrOH). [Chapman et al. *J Chem Soc* 1925 1959, Beilstein 20 III/IV 1499, 20/4 V 100.]

**4-Methylpiperidine** [626-58-4] M 99.2, b 124.4°/755mm,  $d_4^{20}$  0.839,  $n_D^{25}$  1.4430,  $pK^{25}$  10.78. Purify it *via* the *hydrochloride* (m 189°). It is freed from 3-methylpyridine by zone melting. The *hydrobromide* has m 173° (from butanone/\*C<sub>6</sub>H<sub>6</sub>). [Beilstein 20 III/IV 1511, 20/4 V 116.]

**1-Methyl-4-piperidone** [1445-73-4] M 113.2, b 53-56°/0.5mm, 54-56°/9mm, 68-71°/17mm, 85-87°/45mm,  $d_4^{20}$  0.972,  $n_D^{25}$  1.4588,  $pK^{25}$  7.9. It is best purified by fractional distillation. The *hydrochloride* of the hydrate (4-diol) has m 94.7-95.5°, but the anhydrous *hydrochloride* which crystallises from CHCl<sub>3</sub>/Et<sub>2</sub>O has m 165-168° (164-167°), and can also be obtained by sublimation at 120°/2mm. The *oxime* has m 130-132° (from

Me<sub>2</sub>CO). The *methiodide* crystallises from MeOH, the crystals with 1MeOH have **m** 189-190°, and the solvent-free *iodide* has **m** 202-204°(dec). [Lyle et al. *J Org Chem* **24** 342 1959, Bowden & Green *J Chem Soc* 1164 1952, Tomita *Yakugaku Zasshi (J Pharm Soc Japan)* **71** 1053 1951, *Beilstein* **21** III/IV 3183, **21/6** V 419.]

**2-Methylpyrazine** [109-08-0] **M** 94.1, **f** -28.8°, **b** 136-137°, **d**<sub>4</sub><sup>20</sup> 1.025, **n**<sub>D</sub><sup>20</sup> 1.505, **pK**<sub>1</sub><sup>25</sup> -5.25 (aqueous H<sub>2</sub>SO<sub>4</sub>), **pK**<sub>2</sub><sup>25</sup> 1.47. Purify it *via* the picrate and distil the *free base*. The *picrate* has **m** 133-134°(from EtOH). [Wiggins & Wise *J Chem Soc* 4780 1956, *Beilstein* **23** III/IV 911, **23/5** V 386.]

**2-Methylpyridine (2-picoline)** [109-06-8] **M** 93.1, **b** 129.4°, **d**<sub>4</sub><sup>20</sup> 0.9444, **n**<sub>D</sub><sup>20</sup> 1.50102, **pK**<sup>25</sup> 5.96. Biddiscombe and Handley [*J Chem Soc* 1957 1954] steam distilled a boiling solution of the base in 1.2 equivalents of 20% H<sub>2</sub>SO<sub>4</sub> until about 10% of the base had been carried over, along with non-basic impurities. Excess aqueous NaOH is then added to the residue, the free base is separated, dried with solid NaOH and fractionally distilled.

2-Methylpyridine can also be dried with BaO, CaO, CaH<sub>2</sub>, LiAlH<sub>4</sub>, or Linde type 5A molecular sieves. An alternative purification is *via* the ZnCl<sub>2</sub> adduct, which is formed by adding 2-methylpyridine (90ml) to a solution of anhydrous ZnCl<sub>2</sub> (168g) and 42ml conc HCl in absolute EtOH (200ml). Crystals of the complex are filtered off, recrystallised twice from absolute EtOH (to give **m** 118.5-119.5°), and the free base is liberated by addition of excess aqueous NaOH. It is steam distilled, and solid NaOH is added to the distillate to form two layers, the upper one of which is then dried with KOH pellets, stored for several days with BaO and fractionally distilled. Instead of ZnCl<sub>2</sub>, HgCl<sub>2</sub> (430g in 2.4L of hot water) can be used. The complex, which separates on cooling, can be dried at 110° and recrystallised from 1% HCl (to **m** 156-157°). *The alkali metals, Na, Li or Cs should NOT be used for drying pyridine, and pyridine derivatives, as they form coloured pyridine radical anions leading to bipyridyls.* [see Schmulback et al. *J Am Chem Soc* **90** 6600 1968].

The *hydrochloride* has **m** 78-79°, and the *picrate* has **m** 165.5°(from EtOH) and 180°(from H<sub>2</sub>O). [*Beilstein* **20** III/IV 2679, **20/5** V 464.]

**3-Methylpyridine (3-picoline)** [108-99-6] **M** 93.1, **m** -18.5°, **b** 144°/767mm, **d**<sub>4</sub><sup>20</sup> 0.957, **n**<sub>D</sub><sup>20</sup> 1.5069, **pK**<sup>25</sup> 5.70. In general, the same methods of purification that are described for *2-methylpyridine* can be used. However, 3-methylpyridine often contains 4-methylpyridine and 2,6-lutidine, neither of which can be removed satisfactorily by drying and fractionation, or by using the ZnCl<sub>2</sub> complex. Biddiscombe and Handley [*J Chem Soc* 1957 1954], after steam distillation as for *2-methylpyridine*, treated the residue with urea to remove 2,6-lutidine, then azeotropically distilled with acetic acid (the azeotrope had **b** 114.5°/712mm), and recovered the base by adding excess of aqueous 30% NaOH, drying with solid NaOH and carefully fractionally distilling. The distillate is then fractionally crystallised by slow partial freezing. An alternative treatment [Reithoff et al. *Ind Eng Chem (Anal Edn)* **18** 458 1946] is to reflux the crude base (500ml) for 20-24 hours with a mixture of acetic anhydride (125g) and phthalic anhydride (125g) followed by distillation until phthalic anhydride begins to pass over. The distillate is treated with NaOH (250g in 1.5L of water) and then steam distilled. Addition of solid NaOH (250g) to this distillate (*ca* 2L) led to the separation of 3-methylpyridine which is removed, dried (K<sub>2</sub>CO<sub>3</sub>, then BaO) and fractionally distilled. (Subsequent fractional freezing would probably be advantageous.) *The alkali metals, Na, Li or Cs should NOT be used for drying pyridine, and pyridine derivatives, as they form coloured pyridine radical anions leading to bipyridyls.* [see Schmulback et al. *J Am Chem Soc* **90** 6600 1968]. The *hydrochloride* has **m** 85°, and the *picrate* has **m** 153°(from Me<sub>2</sub>CO, EtOH or H<sub>2</sub>O). [*Beilstein* **20** III/IV 2710, **20/5** V 506.]

**4-Methylpyridine (4-picoline)** [108-89-4] **M** 93.1, **m** 4.25°, **b** 145.0°/765mm, **d**<sub>4</sub><sup>20</sup> 0.955, **n**<sub>D</sub><sup>20</sup> 1.5058, **pK**<sup>25</sup> 4.99. It can be purified as for *2-methylpyridine*. Biddiscombe and Handley's method (above) for 3-methylpyridine is also applicable. Lidstone [*J Chem Soc* 242 1940] purified it *via* the *oxalate* (**m** 137-138°) by heating 100ml of 4-methylpyridine to 80° and adding slowly 110g of anhydrous oxalic acid, followed by 150ml of boiling EtOH. After cooling and filtering, the precipitate is washed with a little EtOH, then recrystallised from EtOH, dissolved in the minimum quantity of water and distilled with excess 50% KOH. The distillate is dried with solid KOH and again distilled. Hydrocarbons can be removed from 4-methylpyridine by converting the latter to its hydrochloride, crystallising from EtOH/diethyl ether, regenerating the free base by adding alkali and distilling. As a final purification step, 4-methylpyridine can be fractionally crystallised by partial freezing

to effect a separation from 3-methylpyridine. Contamination with 2,6-lutidine is detected by its strong absorption at 270nm. *The alkali metals, Na, Li or Cs should NOT be used for drying pyridine, and pyridine derivatives, as they form coloured pyridine radical anions leading to bipyridyls.* [see Schmulback et al. *J Am Chem Soc* **90** 6600 1968]. The *hydrochloride* has *m* 161°, and the *picrate* has *m* 167°(from Me<sub>2</sub>CO, EtOH or H<sub>2</sub>O). [*Beilstein* **20** III/IV 2732, **20/5** V 543.]

**2-Methylpyridin-5-yl trifluoromethanesulfonate (2-methyl-5-pyridine triflate)** [111770-91-3] *M* **241.2**, *b* **80-82°/1.9mm**, *d*<sub>4</sub><sup>25</sup> **1.412**, *n*<sub>D</sub><sup>20</sup> **1.442**. Distil the triflate under vacuum, or if it is discoloured dissolve it in CH<sub>2</sub>Cl<sub>2</sub>, wash it with N NaOH and half saturated K<sub>2</sub>CO<sub>3</sub>, dry it over solid K<sub>2</sub>CO<sub>3</sub>, filter, evaporate, and distil by bulb-to-bulb distillation at 65-70°/0.1mm, then redistil *in vacuo*. Its <sup>1</sup>H NMR [(CD<sub>3</sub>)<sub>2</sub>SO] has δ at 2.50 (s, 3H, Me), 7.45 (d, *J* = 9.2 Hz, 1H), 7.90 (dd, *J* = 9.2, 2.3 Hz, 1H) and 8.60 (d, *J* = 2.3 Hz, 1H). [Tilley & Zawoiski *J Org Chem* **53** 386 1988, Ellingboe et al. *J Med Chem* **37** 542 1994.] When stirred with *m*-chloroperbenzoic acid in CH<sub>2</sub>Cl<sub>2</sub> (16 hours, 25°), filtered, concentrated and purified by flash chromatography (2% MeOH/ CH<sub>2</sub>Cl<sub>2</sub>), the triflate gave colourless crystals of the *triflate N-oxide m* **47-48°** with <sup>1</sup>H NMR [(CD<sub>3</sub>)<sub>2</sub>SO] with δ at 2.36 (s, 3H, Me), 7.56 (dd, *J* = 8.9, 2.3 Hz, 1H), 7.69 (d, *J* = 8.9 Hz, 1H) and 8.84 (d, *J* = 2.3 Hz, 1H), Ellingboe et al. *J Med Chem* **37** 542 1994.]

***N*-Methylpyrrole** [96-54-8] *M* **81.1**, *b* **115-116°/756mm**, *d*<sub>4</sub><sup>20</sup> **0.908**, *n*<sub>D</sub><sup>20</sup> **1.487**, *pK*<sup>25</sup> **-3.4 (-2.90)**. Dry *N*-methylpyrrole with CaSO<sub>4</sub>, then fractionally distil it from KOH immediately before use. [*Beilstein* **20** III/IV 2080, **20/5** V 8.]

**1-Methyl-2-pyrrolidinone (1-methyl-2-pyrrolidone)** [872-50-4] *M* **99.1**, *f* **-24.4**, *b* **65-76°/1mm**, **78-79°/12mm**, **94-96°/20mm**, **202°/760mm**, *d*<sub>4</sub><sup>20</sup> **1.0328**, *n*<sub>D</sub><sup>20</sup> **1.4678**, *pK*<sup>25</sup> **-0.17 (also -0.92, and 0.2)**. Dry the pyrrolidone by removing water as the \*benzene azeotrope. Fractionally distil at 10 torr through a 100-cm column packed with glass helices. [Adelman *J Org Chem* **29** 1837 1964, McElvain & Voza *J Am Chem Soc* **71** 896 1949.] The *hydrochloride* has *m* 86-88° (from EtOH or Me<sub>2</sub>CO/EtOH) [Reppe et al. *Justus Liebigs Ann Chem* **596** 1 1955]. [*Beilstein* **21** II 213, **21** III/IV 3145, **21/6** V 321.]

**2-Methylquinoline (quinaldine)** [91-63-4] *M* **143.2**, *b* **86-87°/1mm**, **155°/14mm**, **246-247°/760mm**, *d*<sub>4</sub><sup>20</sup> **1.058**, *n*<sub>D</sub><sup>20</sup> **1.6126**, *pK*<sup>25</sup> **5.65**. Dry it with Na<sub>2</sub>SO<sub>4</sub> or by refluxing with BaO, then fractionally distil it under reduced pressure and redistil it from zinc dust. Purify it further by conversion to its *phosphate* (*m* 220°) or *picrate* (*m* 192°) from which after recrystallisation, the free base is regenerated. [Packer et al. *J Am Chem Soc* **80** 905 1958.] Its ZnCl<sub>2</sub> complex can be used for the same purpose. [*Beilstein* **20** III/IV 3454, **20** V 375.]

**4-Methylquinoline (lepidine)** [491-35-0] *M* **143.2**, *b* **265.5°**, *d*<sub>4</sub><sup>20</sup> **1.084**, *n*<sub>D</sub><sup>20</sup> **1.61995**, *pK*<sup>25</sup> **5.59**. Reflux lepidine with BaO, then fractionally distil it. Further purify it *via* its recrystallised *dichromate salt* (*m* 138°) (from H<sub>2</sub>O). [Cumper et al. *J Chem Soc* 1176 1962.] [*Beilstein* **20** III/IV 3477, **20/7** V 389.]

**6-Methylquinoline** [91-62-3] *M* **143.2**, *b* **258.6°**, *d*<sub>4</sub><sup>20</sup> **1.067**, *n*<sub>D</sub><sup>20</sup> **1.61606**, *pK*<sup>25</sup> **4.92**. Reflux it with BaO, then fractionally distil it. Further purified *via* its recrystallised *ZnCl<sub>2</sub> complex* (*m* 190°). [Cumper et al. *J Chem Soc* 1176 1962, *Beilstein* **20** III/IV 3498, **20/7** V 400.]

**7-Methylquinoline** [612-60-2] *M* **143.2**, *m* **38°**, *b* **255-260°**, *d*<sub>4</sub><sup>20</sup> **1.052**, *n*<sub>D</sub><sup>20</sup> **1.61481**, *pK*<sup>25</sup> **5.29**. Purify it *via* its *dichromate complex* (*m* 149°, after five recrystallisations from water). [Cumper et al. *J Chem Soc* 1176 1962, *Beilstein* **20** III/IV 3497, **20/7** V 402.]

**8-Methylquinoline** [611-32-5] *M* **143.2**, *b* **122.5°/16mm**, **247.8°/760mm**, *d*<sub>4</sub><sup>20</sup> **1.703**, *n*<sub>D</sub><sup>20</sup> **1.61631**, *pK*<sup>25</sup> **4.60**. Purify it as for 2-methylquinoline. The *phosphate* and *picrate* have *m* 158° and *m* 201°, respectively. [*Beilstein* **20** III/IV 3500, **20/7** V 405.]

**(±)-3-Methylsulfolane (3-methyl-tetrahydrothiophene-1,1-dioxide)** [872-93-5] *M* **134.2**, *m* **0.5°**, *b* **101°/2mm**, **125-130°/12mm**, **278-282°/763.5mm**, *d*<sub>4</sub><sup>20</sup> **1.1885**, *n*<sub>D</sub><sup>20</sup> **1.4770**. Distil the sulfolane under vacuum and recrystallise it from Et<sub>2</sub>O at -60° to -70°. An IR film has strong bands at 570 and 500 cm<sup>-1</sup>. [Eigenberger *J Prakt Chem* [2] **131** 289 1931, Freaheller & Katon *Spectrochim Acta* **20** 1099 1964, Whitehead et al. *J Am*

*Chem Soc* **73** 3632 1951, *Beilstein* **17** I 8, **17** III/IV 64.]

(±)-2-Methyltetrahydrofuran [96-47-9] **M 86.1, b 80.0°, d<sub>4</sub><sup>20</sup> 0.856, n<sub>D</sub><sup>20</sup> 1.4053.** Likely impurities are 2-methylfuran, methyl-dihydrofurans and hydroquinone (stabiliser, which is removed by distillation under reduced pressures). It is washed with 10% aqueous NaOH, dried, vacuum distilled from CaH<sub>2</sub>, passed through freshly activated alumina under nitrogen, and refluxed over sodium metal under vacuum. Store it over sodium. [Ling & Kevan *J Phys Chem* **80** 592 1976.] Distil it from sodium under vacuum, and store it with sodium-potassium alloy (this treatment removes water and prevents the formation of peroxides). *Alternatively*, it can be freed from peroxides by treatment with ferrous sulfate and sodium bisulfate, then solid KOH, followed by drying with, and distilling from, sodium, or type 4A molecular sieves under argon. It may be difficult to remove \*benzene if it is present as an impurity (can be readily detected by its ultraviolet absorption in the 249-268nm region). [Ichikawa & Yoshida *J Phys Chem* **88** 3199 1984.] It has also been purified by percolating through Al<sub>2</sub>O<sub>3</sub> and fractionated collecting fraction **b 79.5-80°**. After degassing, the material is distilled onto degassed molecular sieves, then distilled onto anthracene and a sodium mirror. The solvent is then distilled from the green solution onto potassium mirror or sodium-potassium alloy, from which it is distilled again. [Mohammad & Kosower *J Am Chem Soc* **93** 2713 1971.] It should be stored in the presence of 0.1% of hydroquinone or 2,6-di-*tert*-butyl-*p*-cresol as stabiliser. The *R*(+)-*enantiomer* has **b 78-80°/atm** and  $[\alpha]_D^{20} +27.5^\circ$  (neat), and the *S*(-)-*enantiomer* has **b 86°/atm** and  $[\alpha]_D^{20} -27.0^\circ$  (neat) [Iffland & Davis *J Org Chem* **42** 4150 1977, Gagnaire & Butt *Bull Soc Chim Fr* 312 1961, *Beilstein* **17** III/IV 60, **17**/I V 78.] **HARMFUL VAPOURS.**

**3-Methylthiophene** [616-44-4] **M 98.2, b 60°/116mm, 111-113°/atm, 115.5°/atm, d<sub>4</sub><sup>20</sup> 1.024, n<sub>D</sub><sup>20</sup> 1.531.** Dry it with Na<sub>2</sub>SO<sub>4</sub>, then distil it from sodium. [*Beilstein* **17** III/IV 277, **17**/I V 331.]

**6(4)-Methyl-2-thiouracil** [56-04-2] **M 142.2, m 330°(dec), 299-303°(dec), 323-324°(dec), pK<sup>25</sup> 8.1.** Crystallise the thiouracil from a large volume of H<sub>2</sub>O. Purify it further by dissolving in base, adding charcoal, filtering and acidifying with AcOH. Suspend the wet solid (*ca* 100g) in boiling H<sub>2</sub>O (1L), stir and add AcOH (20ml), stir and refrigerate. Collect the product, wash it with cold H<sub>2</sub>O (4 x 200ml), drain it for several hours then place it in an oven at 70° to constant weight. [IR: Short & Thompson *J Chem Soc* 168 1952, Foster & Snyder *Org Synth Coll Vol* **IV** 638 1063, *Beilstein* **24** III/IV 1289.]

**4-Methyl-1,2,4-triazoline-3,5-dione (MTAD)** [13274-43-6] **M 113.1, m 103-104°, m 107-109°.** MTAD is obtained as pink needles by sublimation at 40-50°/0.1mm (see 4-phenyl-1,2,4-triazoline-3,5-dione, PTAD below). [Cookson et al. *Org Synth* **51** 121 1971, Cheng et al. *J Org Chem* **49** 2910 1984, *Beilstein* **26** III/IV 538.]

**2-Methyltricycloquinazoline** [2642-52-6] **M 334.4, m >300°.** Purify it by crystallisation from \*C<sub>6</sub>H<sub>6</sub>, toluene or xylene followed by vacuum sublimation. [*cf.* *Beilstein* **26** III/IV 1932.] **CARCINOGEN.**

**5-Methyltryptamine hydrochloride (3-[2-aminoethyl]-5-methylindole hydrochloride)** [1010-95-3] **M 210.7, m 289-291°(dec), 290-292°, pK<sub>Est(1)</sub> ~ -3 (protonation of ring NH), pK<sub>Est(2)</sub> ~ 9.0 (CH<sub>2</sub>NH<sub>2</sub>), pK<sub>Est(3)</sub> ~ 10.9 (acidic indole NH).** Recrystallise the hydrochloride from H<sub>2</sub>O. The *free base* has **m 96-98°** (from \*C<sub>6</sub>H<sub>6</sub>/cyclohexane) or **m 99-100°** (from petroleum ether), and the *picrate* has **m 243°(dec)** (from EtOH). [Young *J Chem Soc* 3493 1958, Gaddum et al. *Quart J Exp Physiol* **40** 49 1955, Röhm *Hoppe Seyler's Z Physiol Chem* **297** 229 1954, *Beilstein* **22** III/IV 4364, **22**/10 V 167.]

**6-Methyluracil** [626-48-2] **M 126.1, m 270-280°(dec), λ<sub>max</sub> 260nm log<sub>e</sub> 3.97, pK<sub>1</sub> ~1.1, pK<sub>2</sub> 9.8.** Crystallise 6-methyluracil from EtOH or acetic acid. [*Beilstein* **24** III/IV 1281.]

**1-Methyluric acid** [708-79-2] **M 182.1, m >350°, pK<sub>1</sub> 5.75 (basic), pK<sub>2</sub> 10.6 (acidic).** Recrystallise it from H<sub>2</sub>O. Its solubility at 17.5° is 1g in 353ml of H<sub>2</sub>O. [Bergmann & Dikstein *J Am Chem Soc* **77** 691 1955.] It has UV with λ<sub>max</sub> at 231 and 283nm (pH 3), and 217.5 and 292.5nm (pH >12) [Johnson *Biochem J* **5** 133 1952]. [*Beilstein* **26** II 299, **26** III/IV 2621.]

**3-Methyluric acid** [39717-48-1] **M 182.1, m >350°**, **pK<sub>1</sub> 5.75 (6.2), pK<sub>2</sub> >12**. Crystallise it from water. Its solubility at 17.5° is 1g in 19.7L of H<sub>2</sub>O. It has UV with  $\lambda_{\max}$  at 232 and 287 nm (pH 3), and 214 and 292.5nm (pH >12). [Beilstein 26 II 299, 26 III/IV 2621.]

**7-Methyluric acid** [612-37-3; 30409-21-3] **M 182.1, m >380°**, **pK<sub>1</sub> 5.6, pK<sub>2</sub> 10.3**. Crystallise it from water. It has UV with  $\lambda_{\max}$  at 234 and 286nm (pH 3), 237 and 293nm (8.5), and 222 and 296.5nm (pH >12) [Beilstein 26 H 525, 26 II 299, 26 III/IV 2622.]

**9-Methyluric acid** [30345-24-5] **M 182.1, m 385-400°(dec), >400°**. Crystallise it from water. [Beilstein 26 II 299, 26 III/IV 2622.]

**1-Methylxanthine (1-methyl-purin-2,6(3-*H*,7-*H*)-dione)** [6136-37-4] **M 166.1, m >360°** **pK<sub>1</sub><sup>20</sup> 1.3, pK<sub>2</sub><sup>20</sup> 7.9, pK<sub>3</sub><sup>20</sup> 11.8**. Crystallise it from water. It has UV with  $\lambda_{\max}$  at 266nm (pH 2.08), 242.5 and 276nm (pH 9). [Beilstein 26 II 263, 26 III/IV 2329.]

**3-Methylxanthine** [1076-22-8] **M 166.1, m >360°** **pK<sub>1</sub><sup>20</sup> 8.45, pK<sub>2</sub><sup>20</sup> 11.92**. Crystallise it from water. [Beilstein 26 II 263, 26 III/IV 2329.]

**7-Methylxanthine** [552-62-5] **M 166.1, m >380°(dec)** **pK<sub>1</sub><sup>20</sup> 8.42, pK<sub>2</sub><sup>20</sup> >13**. Crystallise it from water. [Beilstein 26 II 263, 26 III/IV 2330.]

**8-Methylxanthine** [17338-96-4] **M 166.1, m 292-293°(dec)**. Crystallise it from water. [Beilstein 26 III/IV 2330.]

**9-Methylxanthine** [1198-33-0] **M 166.1, m 384°(dec)**, **pK<sub>1</sub><sup>20</sup> 2.0, pK<sub>2</sub><sup>20</sup> 6.12, pK<sub>3</sub><sup>20</sup> 10.5 (>13)**. Crystallise it from water. [Beilstein 26 II 263, 26 III/IV 2330.]

**Metrazol (Cardiazol, Leptazol, 3a,4,5,6,7,8-hexahydro-1,2,3,3a-tetraaza-azulene, 1,5-pentamethylene-1,2,3,4-tetrazole)** [54-95-5] **M 138.2, m 61°, b 194°/12mm, pK<sub>Est</sub> ~<0**. Crystallise metrazol from diethyl ether and dry it under vacuum over P<sub>2</sub>O<sub>5</sub>, or distil it. [Schmidt *Chem Ber* 57 704 1924, Beilstein 26 II 213.]

**Morin (hydrate) (2',3,4',5,7-pentahydroxyflavone)** [480-16-0] **M 302.2, m 289-292°, CI 75660, pK<sub>1</sub> 5.3, pK<sub>2</sub> 8.74**. Stir morin at room temperature with ten times its weight of absolute EtOH, then leave overnight to settle. Filter it off, and evaporate under a heat lamp to one-tenth its volume. An equal volume of water is added, and the precipitated morin is filtered off, dissolved in the minimum amount of EtOH and again precipitated with an equal volume of water. The precipitate is filtered off, washed with water and dried at 110° for 1 hour (yield *ca* 2.5%). [Perkins & Kalkwarf *Anal Chem* 28 1989 1956.] It complexes with W and Zr. [Beilstein 18 H 239, 18 III/IV 3468, 18/5 V 492.]

**(-)-Morphine (H<sub>2</sub>O)** [57-27-2] **M 302.2, m 230°(dec), 254°(dec, rapid heating), 260°(Kofler block), [α]<sub>D</sub><sup>23</sup> -130.9° (MeOH), pK<sub>1</sub> 8.31, pK<sub>2</sub> 9.51**. Crystallise the narcotic from MeOH or anisole. It dehydrates at 130°. Its solubility in H<sub>2</sub>O is 0.2g/L at 20° and 0.9g/L at 100°, and in EtOH it is 5g/L at 20° and 10g/L on boiling. The *stypnate* has **m 189°** (from aqueous EtOH). [Beilstein 27 II 118, 27 III/IV 2223.]

**Morpholine** [110-91-8] **M 87.1, f -4.9°, b 128.9°, d<sub>4</sub><sup>20</sup> 1.0007, n<sub>D</sub><sup>20</sup> 1.4540, n<sub>D</sub><sup>25</sup> 1.4533, pK<sup>25</sup> 8.33**. Dry morpholine with KOH, fractionally distil it, then reflux it with Na, and again fractionally distil it. Dermer & Dermer [*J Am Chem Soc* 59 1148 1937] precipitated it as the *oxalate* by adding slowly to slightly more than 1<sup>\*</sup> molar equivalent of oxalic acid in EtOH. The precipitate is filtered off and recrystallised twice from 60% EtOH [1:1 salt has **m 190-195°(dec)**]. Addition of the *oxalate* to concentrated aqueous NaOH regenerated the base, which is separated and dried with solid KOH, then sodium, before being fractionally distilled. The *hydrochloride* has **m 178-179°** (from MeOH/Et<sub>2</sub>O), and the *picrate* has **m 151.6°** (from aqueous EtOH). [Beilstein 27 II 3, 27 III/IV 15.]

§ A polystyrene supported morpholine is commercially available.

**2-(*N*-Morpholino)ethanesulfonic acid (MES)** [4432-31-9] **M 213.3, m >300°(dec), pK<sup>20</sup> 6.15.** Crystallise MES from hot EtOH containing a little water. The *picrate* crystallises from EtOH and has **m 178.8-182°.** [Malkiel & Mason *J Org Chem* **8** 199 1943, *Beilstein* **27** III/IV 370.]

**Murexide (ammonium purpurate)** [3051-09-0] **M 284.2, m >300°, λ<sub>max</sub> 520nm (ε 12,000), pK<sub>2</sub> 9.2, pK<sub>3</sub> 10.9.** The sample may be grossly contaminated with uramil, alloxanthine, etc., and may be difficult to purify. It is better to synthesise it from pure alloxanthine [Davidson *J Am Chem Soc* **58** 1821 1936]. Recrystallise it from water. [Kuhn & Lyman *Chem Ber* **69** 1547 1936, *Beilstein* **25** I 709, **25** III/IV 4236.]

**α-Naphthoflavone (7,8-benzoflavone)** [604-59-1] **M 272.3, m 153-155°, 155°, pK<sup>25</sup> 8-9 (phenolic OH).** Recrystallise the flavone from EtOH or aqueous EtOH. [IR: Cramer & Windel *Chem Ber* **89** 354 1956, UV Pillon & Massicot *Bull Soc Chim Fr* 26 1954, Smith *J Chem Soc* 542 1946, Mahal & Venkataraman *J Chem Soc* 1767 1934.] It is a competitive inhibitor of human estrogen synthase. [Kellis & Vickery *Science* **225** 1032 1984, *Beilstein* **17** III/IV 5550.]

**Naphthol AS-acetate (3-acetoxynaphthoic acid anilide)** [1163-67-3] **M 305.3, m 152°, 160°.** Recrystallise it from hot MeOH and dry *in vacuo* over P<sub>2</sub>O<sub>5</sub>. It has **m 252°** after sublimation at 210-215°. It is slightly soluble in AcOH, EtOH, CHCl<sub>3</sub> or \*C<sub>6</sub>H<sub>6</sub>. It is a fluorogenic substrate for albumin esterase activity. [Chen & Scott *Anal Lett* **17** 857 1984.] At λ<sub>ex</sub> 320nm it has fluorescence at λ<sub>em</sub> 500nm. [Brass & Sommer *Chem Ber* **61** 1000 1928, *Beilstein* **12** II 260, **12** III 960, **12** IV 923.]

**1,5-Naphthyridine** [254-79-5] **M 130.1, m 75°, b 112°/15mm, pK<sup>20</sup> 2.84.** Purify 1,5-naphthyridine by repeated sublimation. The *picrate* crystallises from EtOH with **m 200°(dec).** [UV: Armarego *Physical Methods in Heterocyclic Chemistry* (Ed Katritzky, Academic Press) **Vol III** 133 1971, *Beilstein* **23** II 178, **23** III/IV 1235.]

**1,8-Naphthyridine** [254-60-4] **M 130.1, m 98-99°, pK<sup>20</sup> 3.36.** Purify 1,8-naphthyridine through an Al<sub>2</sub>O<sub>3</sub> column and elute with toluene and petroleum ether, evaporate the eluate, crystallise the residue from petroleum ether (b 60-80°), and sublime it at 80°/13mm. The *picrate* [15936-16-0] has **m 207-208°** (from EtOH), and the *methiodide* has **m 180-181°** (from EtOH). [Hawes & Wibberley *J Chem Soc (C)* 1564 1967, UV: Armarego *Physical Methods in Heterocyclic Chemistry* (Ed Katritzky, Academic Press) **Vol III** 134 1971, *Beilstein* **23** II 178, **23** III/IV 1237.]

**(±)-Naringenin (4',5,7-trihydroxyflavanone)** [480-41-1] **M 272.3, m 251° (phenolic pKs ~8-11).** Crystallise it from EtOH or aqueous EtOH. It has UV with λ<sub>max</sub> at 290nm (EtOH). The *S*(-)-*enantiomer* (natural form) has **m 255-256°** (from EtOH) and [α]<sub>D</sub><sup>20</sup> -28.0° (c 2, EtOH), [α]<sub>D</sub><sup>20</sup> -35.2° (c 1, pyridine). [*Beilstein* **18** H 503, **18** II 164, **18** III/IV 2630.] **Genistein (4',5,7-trihydroxyisoflavone)** [446-72-0] **M 270.2** crystallises from 60% aqueous EtOH or water with **m 297-298°** and [α]<sub>D</sub><sup>20</sup> -28° (c 0.6, 20mM NaOH). [*Beilstein* **18/4** V 594.]

For **Naringin (naringenin 7-rhamnoglucoside)**. See “Carbohydrates” in Chapter 7.

**Neutral Red (2-amino-8-dimethylamino-3-methylphenazine HCl, Basic Red 5, CI 50040)** [553-24-2] **M 288.8, m 290°(dec), pK<sup>25</sup> 6.5.** Crystallise the dye from \*benzene/MeOH (1:1). In aqueous solution it is red at pH 6.8 and yellow at pH 8.0. [*Beilstein* **25** III/IV 3054.]

**Nicotinaldehyde thiosemicarbazone** [3608-75-1] **M 180.2, m 222-223°.** Crystallise the derivative from EtOH, BuOH or water. Its *hydrochloride* crystallises from aqueous EtOH with **m 237-239° (dec).** [*Beilstein* **21** III/IV 3542 **21/7** V 342.]

**Nicotinamide (Niacin, Vitamin B<sub>3</sub>, vitamin PP)** [98-92-0] **M 122.1, m 128-131°, b ~150-160°/high vac, pK<sub>1</sub><sup>20</sup> 0.5, pK<sub>2</sub><sup>20</sup> 3.33.** Crystallise niacin from \*benzene. It has solubility in g/ml: H<sub>2</sub>O (1), EtOH (0.7) and glycerol (0.1). [*Methods in Enzymology* **66** 23 1980, UV: Armarego *Physical Methods in Heterocyclic Chemistry* (Ed Katritzky, Academic Press) **Vol III** 83 1971, *Beilstein* **22** III/IV 389, **22/2** V 80.]

**Nicotinic acid (Niacin is also used for the acid, pyridine-3-carboxylic acid)** [59-67-6] **M 123.1, m 232-234°**, **pK<sub>1</sub><sup>25</sup> 2.00, pK<sub>2</sub><sup>25</sup> 4.82**. Crystallise the acid from \*benzene, EtOH or H<sub>2</sub>O. It sublimes without decomposition. [McElvain *Org Synth Coll Vol I* 385 1941, *Beilstein* **22** III/IV 439, **22/2** V 57.]

**Nicotinic acid hydrazide** [553-53-7] **M 137.1, m 158-159°**, **pK<sub>1</sub><sup>25</sup> 2.2, pK<sub>2</sub><sup>25</sup> 3.63, pK<sub>3</sub><sup>25</sup> 11.49(NH)**. Crystallise it from aqueous EtOH or \*benzene. [*Beilstein* **22** III/IV 439, **22/2** V 121.]

**Nile Blue A (a benzophenoxazinium sulfate dye)** [3625-57-8] **M 415.5, m >300°(dec)**, **CI 51180, pK<sup>25</sup> 2.4**. Crystallise the dye from aqueous AcOH. It has UV with  $\lambda_{\max}$  at 630nm (96% aqueous EtOH) and 635nm (H<sub>2</sub>O). The *betaine* has UV with  $\lambda_{\max}$  at 513nm (EtOH). [Crossley et al. *J Am Chem Soc* **74** 57, 578 1952, Merill & Spencer *J Am Chem Soc* **70** 3683 1948, *Beilstein* **27** II 457, **27** III/IV 5166.]

**5-Nitrobarbituric acid (dilituric acid)** [480-68-2] **M 173.1, m 176°, 176-183°(dec)**, **pK<sup>20</sup> 10.25**. Crystallise dilituric acid from water as the *trihydrate* (**m 180-181°** (dec)). Drying over 70% H<sub>2</sub>SO<sub>4</sub> converts the trihydrate to the *dihydrate* [Loeffler & Moore *J Am Chem Soc* **70** 3650 1948]. [*Beilstein* **24** H 474, **24** II 273, **24** III/IV 1882.]

**4'-Nitrobenzo-15-crown-5** [60835-69-0] **M 313.3, m 84-85°, 93-95°**. Recrystallise the crown ether from EtOH, MeOH or \*C<sub>6</sub>H<sub>6</sub>/hexane as for the 18-crown-6 compound below. It complexes with Na<sup>+</sup>, K<sup>+</sup>, NH<sub>4</sub><sup>+</sup>, Ca<sup>2+</sup>, Mg<sup>2+</sup> and Cd<sup>2+</sup>. The <sup>1</sup>HNMR spectrum (CDCl<sub>3</sub>) has  $\delta$  (ppm) at 3.6-4.4 (m 16CH<sub>2</sub>), 6.8 (d 1H arom), 7.65 (d 1H arom), 7.80 (dd 1H arom  $J_{ab} = 9\text{Hz}$  and  $J_{bc} = 3\text{Hz}$ ) [Schmid et al. *J Am Chem Soc* **98** 5198 1976, Kikukawa et al. *Bull Chem Soc Jpn* **50** 2207 1977, Toke et al. *Justus Liebigs Ann Chem* **349** 349, 761 1988, Lindner et al. *Z Anal Chem* **322** 157 1985].

**4'-Nitrobenzo-18-crown-6** [53408-96-1] **M 357.4, m 83-84°, 83-84°**. If impure and discoloured, then chromatograph it through Al<sub>2</sub>O<sub>3</sub> and elute with \*C<sub>6</sub>H<sub>6</sub>/hexane (1:1) containing 1% MeOH. The fractions are followed by TLC on Al<sub>2</sub>O<sub>3</sub> (with Dragendorff's reagent for detection: R<sub>F</sub> 0.6 in the above solvent system). Recrystallise the residues from the required fractions from \*C<sub>6</sub>H<sub>6</sub>/hexane to give yellowish leaflets. It complexes with Na or K ions with logK<sub>Na</sub> 3.95 and logK<sub>K</sub> 4.71. [Petranek & Ryba *Col Chem Czech Chem Commun* **39** 2033 1974.]

**4-(4-Nitrobenzyl)pyridine (PNBP)** [1083-48-3] **M 214.2, m 70-71°, 74°, pK<sub>Est</sub> ~5.0**. Crystallise PNBP from aqueous EtOH or cyclohexane. The *hydrochloride* has **m 194-196°(dec)**, from EtOH, and the *picrate* has **m 168°(dec)**, from EtOH/EtOAc. [*Beilstein* **20** II 272, **20/7** V 564.]

**5-Nitroindole** [6146-52-7] **M 162.1, m 141-142°, pK<sup>25</sup> -7.4 (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Decolourise (charcoal) 5-nitroindole and recrystallise it twice from aqueous EtOH or recrystallise it from octane. It has UV has  $\lambda_{\max}$  at 265 and 324nm (EtOH). [*Beilstein* **20** III/IV 3194, **20/7** V 41.]

**Nitron [1,4-diphenyl-3-phenylamino-(1H)-1,2,4-triazolium (hydroxide) inner salt]** [2218-94-2] **M 312.4, m 189°(dec)**. Crystallise it from EtOH, chloroform or EtOH/\*C<sub>6</sub>H<sub>6</sub>. [*Beilstein* **25** III/IV 1075.]

**5-Nitro-1,10-phenanthroline** [4199-88-6] **M 225.2, m 197-198°, pK<sup>25</sup> 3.33**. Crystallise the phenanthroline from \*benzene/petroleum ether, until anhydrous. It also crystallises from H<sub>2</sub>O with **m 202°**, and EtOAc with **m 203°**. Its pK<sup>25</sup> varies from 3.20 to 2.69 with varying MeOH/H<sub>2</sub>O ratios from 0 to 0.95 moles/L, and from 3.20 to 1.95 in varying EtOH/H<sub>2</sub>O ratios from 0 to 0.94 moles/L [Ram et al. *J Prakt Chem* **319** 719 1977]. It forms complexes with Cu<sup>2+</sup>, Zn<sup>2+</sup>, In<sup>2+</sup>, Fe<sup>2+</sup>, Co<sup>2+</sup>, Ni<sup>2+</sup>. [*Beilstein* **23** III/IV 1682, **23/8** V 425.]

**3-Nitro-2-pyridinesulfonyl chloride** [68206-45-1] **M 190.2, m 217-222°(dec)**. The chloride crystallises as yellow needles from CH<sub>2</sub>Cl<sub>2</sub>. When pure, it is stable for several weeks at room temperature, and no decomposition was observed after 6 months at <0°. It is moisture sensitive. Its UV (MeCN) has  $\lambda_{\max}$  at 231nm ( $\epsilon$  12,988), 264nm ( $\epsilon$  5,784) and 372nm ( $\epsilon$  3,117). [NMR and UV: Matsuda & Aiba *Chem Lett* 951 1978, Wagner et al. *Chem Ber* **75** 935 1942.]

**5-Nitroquinoline** [607-34-1] M 174.2, m 70°, pK<sup>20</sup> 2.69. Crystallise 5-nitroquinoline from pentane, then from \*benzene. The *hydrochloride* has m 224° and the *picrate* has m 206°, 214°(from MeOH). [Beilstein 20 H 371, 20 II 235, 20 III/IV 3397.]

**8-Nitroquinoline** [706-35-2] M 174.2, m 88-89°, 91-92°, pK<sup>20</sup> 2.55. Crystallise 8-nitroquinoline from hot water, MeOH, EtOH or EtOH/diethyl ether (3:1). It sublimes at 70°/2mm. [Beilstein 20 H 373, 20 III/IV 3399.]

**4-Nitroquinoline 1-oxide** [56-57-5] M 190.2, m 154-155°, 157°. The *N*-oxide recrystallises from aqueous acetone as yellow needles or platelets. [Ochiai *J Org Chem* 18 534 1953, Seki et al. *J Phys Chem* 91 126 1987, Beilstein 20 III/IV 3396.]

**5-Nitouracil (2,4-dihydroxy-5-nitropyrimidine)** [611-08-5] M 157.1, m 280-285°, >300°, pK<sup>20</sup><sub>1</sub> 0.03, pK<sup>20</sup><sub>2</sub> 5.55, pK<sup>20</sup><sub>3</sub> 11.3. The uracil recrystallises in prisms from boiling H<sub>2</sub>O as the *monohydrate* and loses H<sub>2</sub>O on drying *in vacuo*. [UV: Brown *J Chem Soc* 3647 1959, Brown *J Appl Chem* 2 239 1952, Johnson *J Am Chem Soc* 63 263 1941, Beilstein 24 I 313, 24 II 171, 24 III/IV 1236.]

**4-Nonadecylpyridine (hydrogen ionophore II [ETH 1907] - Proton ionophore)** [70268-36-9] M 345.6, b 180°/0.07mm, pK<sub>Est</sub> ~6.0. Dissolve the waxy ionophore (ca 60g) in CHCl<sub>3</sub> (200ml), wash it with H<sub>2</sub>O (3 x 200ml), dry it and evaporate it to dryness, then distil it in a vacuum. A waxy solid is formed on cooling the distillate. Its UV has λ<sub>max</sub> at 257nm (ε 1.86 x 10<sup>3</sup> M<sup>-1</sup>cm<sup>-1</sup>), 308nm (ε 1.7 x 10<sup>2</sup> M<sup>-1</sup>cm<sup>-1</sup>). [IR, NMR UV: Valenty et al. *Inorg Chem* 18 2160 1979.]

**Norcodeine** [467-15-2] M 285.3, m 185°, 186°, pK<sup>20</sup> 9.10. It crystallises from acetone or ethyl acetate. [Speyer & Walther *Chem Ber* 63 822 1930.] The *hydrochloride* has m 309°(dec) when crystallised from H<sub>2</sub>O. [Beilstein 18 III/IV 8091.]

**Octadecyl isonicotinate (hydrogen ionophore IV ETH 1778)** [103225-02-1] M 375.6, m 57.5°, pK<sub>Est</sub> ~3.5. Dissolve it in Et<sub>2</sub>O and wash it 3 times with H<sub>2</sub>O. Dry the extract (MgSO<sub>4</sub>), evaporate, and recrystallise the residue from EtOAc/hexane (4:1). [Oesch et al. *Anal Chem* 58 2285 1986.]

**Orotic acid (H<sub>2</sub>O)** [50887-69-9] M 174.1, m 334°(dec), 235-346°(dec), pK<sup>25</sup><sub>1</sub> 1.8, pK<sup>25</sup><sub>2</sub> 9.55. It crystallises from water. The *anhydrous acid* [65-86-1] M 156.1 has m 245-246°(dec). [Nye & Mitchell *J Am Chem Soc* 69 1382 1947, Beilstein 25 III/IV 1759.]

**Orotic acid Li salt H<sub>2</sub>O (1-carboxy-4,6-dihydroxypyrimidine Li salt H<sub>2</sub>O)** [5266-20-6] M 180.0, m >300°, pK<sub>1</sub> 2.8 (CO<sub>2</sub>H), pK<sub>2</sub> 9.4 (OH), pK<sub>3</sub> >13 (OH) (for free acid). The salt is soluble in H<sub>2</sub>O at 17° and 100°. Best to acidify an aqueous solution of the salt, isolate the *free acid* (see above) which is recrystallised from H<sub>2</sub>O (as *monohydrate*) m 345-347° (345-346°), then dissolve it in EtOH, add an equivalent amount of LiOH in EtOH and evaporate. Its solubility in H<sub>2</sub>O is 1.28% (17°) and 2.34% (100°). [Bachstsz *Chem Ber* 63 1000 1930, Johnson & Shroeder *J Am Chem Soc* 54 2941 1932, UV: Shugar & Fox *Biochim Biophys Acta* 9 199 1952, Beilstein 25 III/IV 1759.]

**Oxalylindigo** [2533-00-8] M 316.3. It crystallises twice from nitrobenzene as small yellow crystals and is dried by heating *in vacuo* for several hours. [Schanze et al. *J Am Chem Soc* 108 2646 1986.]

**2-Oxazolidinone (ethylene carbamate)** [497-25-6] M 87.1, m 88-90°, 89-90°, 91°, b 152°/0.4mm, 200°/12mm. It is prepared by reaction of ethanolamine with phosgene or diethylcarbonate. It can be prepared from ethanolamine (2g) in CHCl<sub>3</sub> (200ml, EtOH free by passing through an Al<sub>2</sub>O<sub>3</sub> column) by bubbling COCl<sub>2</sub> through the solution which is allowed to stand for 3 hours, the acid is neutralised with powdered PbCO<sub>3</sub>, filtered, evaporated to dryness and the solid residue is recrystallised from CHCl<sub>3</sub>. *Alternatively*, ethanolamine (61g), Et<sub>2</sub>CO<sub>3</sub> (150ml) and NaOMe (0.5g) are heated in an oil bath; and after the EtOH (~112ml) has distilled off, the residue solidifies on cooling and is recrystallised from CHCl<sub>3</sub> (100ml) to give the *oxazolidinone* (57g).

65%), m 87-89°. Its IR (film) has  $\nu_{\max}$  at 3000, 2920 (CH), 1766 (carbamate, C=O), 1690 (amide, C=O), 1465, 1382 and 1357 (CH<sub>2</sub>), 1298 (C-N), 1212, 1140 and 1030 (COC, C-O), 953 cm<sup>-1</sup>; and <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 6.68 (NH, br s, 1H), 4.47 (H<sub>5</sub>, t, <sup>3</sup>J = 8Hz, 2H), 3.63 (H<sub>4</sub>, t, <sup>3</sup>J = 8Hz, 2H); and <sup>13</sup>C NMR (CDCl<sub>3</sub>) with  $\delta$  at 161.2 (C<sub>2</sub>), 65.1 (C<sub>5</sub>), 40.8 (C<sub>4</sub>); and <sup>13</sup>C NMR (D<sub>2</sub>O) with  $\delta$  at 165.1 (C<sub>2</sub>), 68.8 (C<sub>5</sub>), 43.1 (C<sub>4</sub>). [Hammer et al. *J Org Chem* **46** 1521 1981.] [Homeyer US Pat 2,399,118 1964, *Chem Abs* **40** 4084 1964.] It can be recrystallised from \*benzene, dichloroethane, CHCl<sub>3</sub> or EtOH. It is a cyclic urethane, and as such it is not very stable in aqueous solvents. The *N*-acetyl derivative [1432-43-5] **M 129.1, m 69-70°**, is obtained by boiling 2-oxazolidinone (3g) with Ac<sub>2</sub>O (20ml) and NaOAc (1g) for 1.5 hours, and is recrystallised from \*C<sub>6</sub>H<sub>6</sub>/Et<sub>2</sub>O or sublimed at 65° *in vacuo*. [Homeyer US Pat 2,399,118 1964, *Chem Abs* **40** 4084 1964.] The *N*-methyl derivative [19836-78-3] **M 101.1, b 120°/0.1mm**, is prepared from *N*-methyl ethanolamine with COCl<sub>2</sub>/CHCl<sub>3</sub>/PbCO<sub>3</sub> or as for the preparation of the parent compound [with (EtO)<sub>2</sub>CO], and distilled in a vacuum. It is **TOXIC**. The yellow *picrolonate* has **m 137-138°**. [*cf.* Fränkel & Cornelius *Chem Ber* **51** 1662 1918, Ben-Ishai *J Am Chem Soc* **78** 4962 1956, *Beilstein* **27** H 135, **27** I 259, **27** III/IV 2516.]

**Oxetane (1,3-trimethylene oxide)** [503-30-0] **M 58.1, b 45-46°/736mm, 47-49°/atm, 48°/760mm, d<sub>4</sub><sup>20</sup> 0.892, n<sub>D</sub><sup>20</sup> 1.395**. Distil oxetane twice from sodium metal and then fractionate it through a small column at atmospheric pressure, **b 47.0-47.2°**. It can also be purified by preparative gas chromatography using a 2m silica gel column. *Alternatively*, add KOH pellets (50g for 100g of oxetane) and distil it through an efficient column or a column packed with 1/4in Berl Saddles with the main portion boiling at 45-50° being collected and redistilled over fused KOH. [Noller *Org Synth Coll Vol III* 835 1955, Dittmer et al. *J Am Chem Soc* **79** 4431 1957, *Beilstein* **17** H 6, **17** I 3, **17** II 12, **17** III/IV 13, **17** I V 11.]

**Oxetan-2-one (β-propiolactone, propan-3-olide)** [57-57-8] **M 72.1, m -31.2°, -35°, b 51°/10mm, 83°/45mm, d<sub>4</sub><sup>20</sup> 1.1460, n<sub>D</sub><sup>25</sup> 1.4117**. Fractionally distil the lactone from sodium under reduced pressure. It gives an acidic solution in H<sub>2</sub>O. It irritates the skin and is a possible **carcinogen**. [*Beilstein* **17** I 130, **17** III/IV 4157.]

**Oxine Blue [3-(4-hydroxyphenyl)-3-(8-hydroxy-6-quiniliny)-1(3H)-isobenzofuranone]** [3733-85-5] **M 369.4, m 134-135°**. Recrystallise the dye from EtOH and dry it in a desiccator over H<sub>2</sub>SO<sub>4</sub>.

**Oxolinic acid (5-ethyl-5,8-dihydro-8-oxo-1,3-dioxolo[4,5-g]quinoline-3-carboxylic acid)** [14698-29-4] **M 261.2, m 313-314°(dec), 314-316°(dec), pK<sub>Est</sub> ~ 2.3**. Purify the acid by recrystallisation from aqueous Me<sub>2</sub>CO, 95% EtOH or dimethylformamide. It has UV with  $\lambda_{\max}$  at 220, (255.5sh), 259.5, 268, (298sh, 311sh), 321 and 326nm [ $\epsilon$  14.8, (36.8sh), 38.4, 38.4, (6.4sh, 9.2sh), 10.8 and 11.2 x 10<sup>3</sup>]. [Kaminsky & Metzger *J Med Chem* **11** 160 1968, *Beilstein* **17** H 6, **17** I 3, **17** II 202, **17** III/IV 13, **17** I V 11.]

**Papaverine hydrochloride (6,7-dimethoxy-1-veratrylisoquinoline hydrochloride)** [61-25-6] **M 375.9, m 215-220°, 222.5-223.5°(dec), 231°, pK<sub>25</sub> 6.41**. Recrystallise it from H<sub>2</sub>O. It sublimes at 140°/0.1mm. Its solubility in H<sub>2</sub>O is 5%. [Saunders & Srivastava *J Pharm Pharmacol* **3** 78 1951, Biggs *Trans Faraday Soc* **50** 800 1954.] The *free base* has **m 148-150°**. The *picrate* has **m 186-189°(dec), 186-186.5°(dec)** [Bobbitt *J Org Chem* **22** 1729 1957]. [*Beilstein* **21** II 202, **21** III/IV 2788, **21**/6 V 182.]

**Paraldehyde (acetaldehyde trimer, 2r,4c,6c-trimethyl-1,3,5-trioxane, all-cis)** [123-63-7] **M 132.2, m 12.5°, b 124°/751mm, d<sub>4</sub><sup>20</sup> 0.995, n<sub>D</sub><sup>20</sup> 1.407**. Wash paraldehyde with water and fractionally distil it. *Alternatively*, it is purified by drying with anhydrous Na<sub>2</sub>SO<sub>4</sub>, then cooled to 5°, and the frozen material is separated by decantation. The solid is distilled (**b 121-124°/atm**), the distillate is collected, stored over anhydrous Na<sub>2</sub>SO<sub>4</sub> for several days and re-distilled at atmospheric pressure before use [Le Fevre et al. *J Chem Soc* 290 1950]. The *2r,4c,6t-trimethyl-1,3,5-trioxane* has **m 14.5°, b 125°/760mm**. [*Beilstein* **19** II 394, **19** III/IV 4715, **19**/9 V 112.]

**Patulin [4-hydroxy-4H-furo(3,2-c)pyran-2(6H)-one, Clavatin]** [149-29-1] **M 154.1, m 110°, 111-112°**. Crystallise patulin from \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O, EtOH or chloroform. It sublimes at 90°/high vacuum [Bergel et al. *J Chem Soc* 415 1944]. [*Beilstein* **18** III/IV 1184, **18**/3 V 5.] (*Highly TOXIC*).

**Pentachloropyridine** [2176-62-7] **M** 251.3, **m** 122-124°, 123°, 124°, 124-125°, 125-126°, **b** 279-280°/atm, **pK**<sup>20</sup> -6.02 (aqueous H<sub>2</sub>SO<sub>4</sub>). Purify it by recrystallisation from EtOH or aqueous EtOH. It sublimes at 150°/3mm. [den Hertog et al. *Rec Trav Chim Pays Bas* **69** 673 1950, Schikh et al. *Chem Ber* **69** 2604 1936, *Beilstein* **20** I 81, **20** III/IV 2503, **20/5** V 422.]

**Pentafluoropyridine** [700-16-3] **M** 169.1, **m** -41.5°, **b** 83.5°, 84°, 83-85°, **d**<sub>4</sub><sup>20</sup> 1.609, **n**<sub>D</sub><sup>20</sup> 1.3818, **pK**<sub>Est</sub> <0. Distil it through a concentric tube column; it has λ<sub>max</sub> in cyclohexane at 256.8nm. [Chambers et al. *J Chem Soc* 3573 1964, <sup>19</sup>F NMR: Bell et al. *J Fluorine Chem* **1** 51 1971.] The *hexafluoroantimonate* has **m** 98-102°(dec) after crystallisation from liquid SO<sub>2</sub>. [*Beilstein* **20/5** V 401.]

**3,3,6,9,9-Pentamethyl-2,10-diazabicyclo[4.4.0]dec-1-ene (PMDBD, 1,2,3,4,4a,5,6,7-octahydro-2,2,4a,7,7-pentamethylnaphthyridine)** [69340-58-5] **M** 208.3, **m** 15°, **b** 65°/0.07mm, **d**<sub>4</sub><sup>20</sup> 0.924, **n**<sub>D</sub><sup>4</sup> 1.4840, **pK** >11 (MeOCH<sub>2</sub>CH<sub>2</sub>OH/H<sub>2</sub>O, 1:1). The pentamethylhydronaphthyridine is obtained by intramolecular cyclisation of 5-amidino-2,5,8-trimethylnona-2-7-diene [190.0g, 0.914mol, prepared from 2mols BuLi, one mol of propionitrile and 2mols of 3-methylbut-2-enyl bromide (isoprenehydrobromide) in THF at -78°, followed by NaNH<sub>2</sub> in boiling \*C<sub>6</sub>H<sub>6</sub>] in CH<sub>2</sub>Cl<sub>2</sub> (500ml) solution at 0°, by bubbling dry HCl gas through it until the pH is 4. The solution is evaporated *in vacuo* until free from HCl vapour to give a dark hygroscopic hydrochloride which is heated in a flask (protected from air with a CaCl<sub>2</sub> tube) at 200°, whereby excess of HCl is released after 30 minutes (cyclisation begins at *ca* 125°), and heating is continued for 18 hours. The dark brown residue is treated with H<sub>2</sub>O (1.0L), acidified with 2N H<sub>2</sub>SO<sub>4</sub> (500ml) and the aqueous phase is extracted with Et<sub>2</sub>O (2 x 100ml), the aqueous phase is basified (cooling is necessary) with NaOH (*ca* 100g) to pH >10. The brown oil that separates is extracted with Et<sub>2</sub>O (2 x 500ml), dried over K<sub>2</sub>CO<sub>3</sub>, filtered, and evaporated to give a viscous oil (190g) which is distilled through a Vigreux column at 0.05mm, and after a small fore run (3.8g), a colourless viscous oil of the *naphthyridine* distils (156.6g, 80% yield, **b** 60-67°). An analytical sample has **m** ~15° and **b** 65°/0.07mm. Store it in an inert dry atmosphere as it absorbs CO<sub>2</sub> in the presence of moisture to form a *bicarbonate* (see below). Its UV/VIS has λ<sub>max</sub> (ε) at 219 (11,600) nm; the IR (7.5% in CHCl<sub>3</sub>) has ν<sub>max</sub> at 3660w, 3350w, 2970s, 2930s, 2860s, 2500w, 1630s, 1475m/sh, 1460m, 1415w, 1385s, 1375m, 1365m, 1360m, 1335m, 1325m, 1250m, 1235m, 1180m, 1150s, 1130m, 1065w, 1000w, 875w, 840w, 655m, 630w, with the immonium band of the hydrochloride in CHCl<sub>3</sub> at 2500 cm<sup>-1</sup>; the <sup>1</sup>H NMR (CCl<sub>4</sub>) of the free base has δ at 1.08-1.13 (2s, 12H, 4CH<sub>3</sub>), 1.19 (s, 3H, CH<sub>3</sub>), 1.2-2.0 (m, 8H), 3.38 (s, NH); MS (200°) has *m/z* at 209 (4), 208 (*M*<sup>+</sup> 25), 194 (25), 193 (100).

The *hydrochloride*, when recrystallised twice from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O (15 hours at 0°), has **m** 195-196°, and the <sup>1</sup>H NMR (CCl<sub>4</sub>) has δ at 1.34-1.13 (2s, 9H, 3CH<sub>3</sub>), 1.42 (s, 6H, 2CH<sub>3</sub>), 1.6-2.3 (m, 8H), 9.78 (br s, 2H NH).

The *hydrogen carbonate salt* is formed from the free base and moist CO<sub>2</sub> in 2 minutes, then by extraction with CH<sub>2</sub>Cl<sub>2</sub>, drying (K<sub>2</sub>CO<sub>3</sub>) and evaporating gives a labile salt decomposing at 122-125°, and the <sup>1</sup>H NMR has δ at 1.28 (s, 6H, 2CH<sub>3</sub>), 1.30 (s, 3H, CH<sub>3</sub>), 1.36 (s, 6H, 2CH<sub>3</sub>), 1.4-2.2 (m, 8H), and **no** NH/OH signals except in CHCl<sub>3</sub> where an intense singlet appears at 7.25.

The *benzoate* crystallises from CH<sub>2</sub>Cl<sub>2</sub> with **m** 183-185°, the *hydrogen sulfate* crystallises from CH<sub>2</sub>Cl<sub>2</sub>/hexane with **m** 208-209°, the *tosylate* has **m** 114°, and the *dihydrogen phosphate* crystallises from MeOH/Et<sub>2</sub>O with **m** 225-228°. [Heinzer, Soukup and Eschenmoser *Helv Chim Acta* **61** 2851 1978, Sternbach et al. *Angew Chem* **91** 670 1979, *Beilstein* **23/5** V 302.] The strong basicity, proximity of the nitrogen atoms, and the steric hinderance from methyl groups around them, confer remarkable coordination properties to this amidine without the usual nucleophilic behaviour of nitrogen bases [Eschenmoser et al. *Angew Chem* **91** 670 1979, Boyle et al. *J Chem Soc, Chem Commun* 239 1992, *Denmark J Org Chem* **46** 3144 1981].

**Pentaquine monophosphate (1,4-pentanediamine, *n*-[6-methoxy-8-quinolinyl]-*N'*-[1-methylethyl] (1:1) phosphate)** [5428-64-8] **M** 399.4, **m** 189-190°, **pK**<sup>70</sup> 8.22. Crystallise it from H<sub>2</sub>O or 95% EtOH. The *free base* has **b** 165-170°/0.05mm, **n**<sub>D</sub><sup>25</sup> 1.5785. The *picrate* crystallises from Me<sub>2</sub>CO/EtOH with **m** 164.5-165.5°. [Drake et al. *J Am Chem Soc* **68** 1529 1946, *Beilstein* **22** III/IV 5814.]

**(±)-Pentobarbital (5-ethyl-5-1'-methylbutyl barbituric acid, Nembutal is the Na salt)** [76-74-4] **M** 226.4, **m** ~127°(dec), **pK**<sub>Est(1)</sub> ~8.0, **pK**<sub>Est(2)</sub> ~12.7. A solution of the sodium salt in 10% HCl is prepared, and the acid is extracted with ether. Evaporation of the extract gives a solid which is then purified by repeated crystallisation

from  $\text{CHCl}_3$ . It sublimes at  $95\text{--}105^\circ/10\text{--}12\text{mm}$ . [Bucket & Sandorfy *J Phys Chem* **88** 3274 1984.] The (+)- and (-)-*enantiomers* crystallise from 50% aqueous EtOH with *m*  $120\text{--}121^\circ$  and have  $[\alpha]_D^{25} +4.73^\circ$  to  $-4.93^\circ$  (EtOH) [Kleiderer & Shonle *J Am Chem Soc* **56** 1772 1934]. [Beilstein **24** I 419, **24** II 287, **24** III/IV 1951.]

**Pericyazine** [10-{3-(4-hydroxy-1-piperidinyl)-propyl}-10*H*-phenothiazine-2-carbonitrile, **Propericiazine**] [2622-26-6] *M* **365.4**, *m* **116-117°**. It recrystallises from a saturated solution in cyclohexane. It is antipsychotic and is a sensitive reagent for Pd, Va, Ru, Rh and Au. [Gowda et al. *Anal Chem* **55** 1816 1983, *Anal Chim Acta* **154** 347 1983, *Beilstein* **27** III/IV 4110.]

**Phenanthridine** (benzo[*c*]quinoline, 3,4-benzoquinoline) [229-87-8] *M* **179.2**, *m* **106.5°**, **108-109°**, **b** **350°/760mm**, **pK<sup>20</sup> 4.61 (4.48)**. Purify it *via* the  $\text{HgCl}_2$  addition compound formed when phenanthridine (20g) in 1:1 HCl (100ml) is added to aqueous  $\text{HgCl}_2$  (60g in 3L), and the mixture is heated to boiling. The  $\text{HgCl}_2$  complex separates as yellow red crystals with *m*  $195\text{--}198^\circ$  [Arcus & Mesley *J Chem Soc* 1780 1953]. Conc HCl is then added until all of the solid has dissolved. The compound separates on cooling and is decomposed with aqueous NaOH (*ca* 5M). Phenanthridine is extracted into  $\text{Et}_2\text{O}$ , evaporated, and the residue is crystallised from petroleum ether (*b*  $80\text{--}100^\circ$ ) or EtOAc. [Cumper et al. *J Chem Soc* 45218 1962.] It is also purified by chromatography on activated alumina from \*benzene solution, with diethyl ether as eluent. Evaporation of ether gives crystalline material which is freed from residual solvent under vacuum, then further purified by fractional crystallisation, under  $\text{N}_2$ , from its melt. It was purified by zone melting and sublimes in a vacuum. The *picrate* has *m*  $218.5\text{--}219.5^\circ$  (from *iso*-PrOH) (also reported are *m*  $244\text{--}245^\circ$  and  $247\text{--}248^\circ$ , from EtOH or  $\text{H}_2\text{O}$ ). [Slough & Ubbelohde *J Chem Soc* 911 1957.] [Beilstein **20** H 466, **20** III/IV 4016, **20/8** V 223.]

**1,10-Phenanthroline** (*o*-phenanthroline) [66-71-7 (*anhydr*); 5144-89-8 ( $\text{H}_2\text{O}$ )] *M* **198.2**, *m* **98-101°**, **108-110° (hydrate)**, **118° (anhydrous)**, **b** **>300°**, **pK<sub>1</sub><sup>25</sup> -0.7 (aqueous  $\text{HClO}_4$ )**, **pK<sub>2</sub><sup>25</sup> 4.86 (4.96)**. Crystallise its *picrate* (*m*  $191^\circ$ ) from EtOH; then the free base is liberated with aqueous alkali, dried at  $78^\circ/8\text{mm}$  over  $\text{P}_2\text{O}_5$  and crystallised from petroleum ether (*b*  $80\text{--}100^\circ$ ). [Cumper et al. *J Chem Soc* 1188 1962.] It can be purified by zone melting. It has also been crystallised from hexane, \*benzene/petroleum ether (*b*  $40\text{--}60^\circ$ ) or sodium-dried \*benzene, dried and stored over  $\text{H}_2\text{SO}_4$ . The *monohydrate* is obtained by crystallisation from aqueous EtOH or ethyl acetate. It has been crystallised from  $\text{H}_2\text{O}$  (300 parts) to give the *monohydrate* *m*  $102\text{--}103^\circ$  which sublimes at  $10^{-3}\text{mm}$  [Fielding & LeFevre *J Chem Soc* 1811 1951.] The *anhydrous* compound has *m*  $118^\circ$  (after drying at high vacuum at  $80^\circ$ ) and is also obtained by recrystallisation from petroleum ether or \* $\text{C}_6\text{H}_6$  (70 parts) and drying at  $78^\circ/8\text{mm}$ . [UV: Badger et al. *J Chem Soc* 3199 1951.] It has a pKa in  $\text{H}_2\text{O}$  of 4.857 ( $25^\circ$ ) or 5.02 ( $20^\circ$ ) and 4.27 in 50% aqueous EtOH ( $20^\circ$ ). [Albert et al. *J Chem Soc* 2240 1948]. [Beilstein **23** H 227, **23** II 235, **23/8** V 419.]

**1,10-Phenanthroline hydrochloride** (*o*-phenanthroline hydrochloride) [3829-86-5] *M* **243.7**, *m* **212-219°**. The hydrochloride crystallises from 95% EtOH, *m*  $212\text{--}219^\circ$  as the *monohydrate*; the *half hydrate* has *m*  $217^\circ$ . The *3HCl* has *m*  $143\text{--}145^\circ$  (sinters at  $128^\circ$ ). [Thevenet et al. *Acta Cryst Sect B* **33** 2526 1977]. [Beilstein **23** II 235, **23/8** V 421.]

**4,7-Phenanthroline-5,6-dione** [84-12-8] *M* **210.2**, *m* **295°(dec)**. The dione crystallises from MeOH. The *mono-oxime* forms yellow crystals from MeOH with *m*  $250^\circ(\text{dec})$ , the *di-oxime* forms yellow crystals from MeOH with *m*  $300^\circ(\text{dec})$  and the *mono-semihydrazone* forms yellow crystals from MeOH with *m*  $195^\circ(\text{dec})$ . [Druey & Schmidt *Helv Chim Acta* **33** 1080 1950, *Beilstein* **24** III/IV 1741.]

**Phenazine** [92-82-0] *M* **180.2**, *m* **171°**, **pK<sub>1</sub><sup>20</sup> -4.9 (aqueous  $\text{H}_2\text{SO}_4$ )**, **pK<sub>2</sub><sup>20</sup> 1.21**. Phenazine crystallises from EtOH,  $\text{CHCl}_3$  or ethyl acetate, after pre-treatment with activated charcoal. It can be sublimed *in vacuo* and purified by zone refining. [Beilstein **23/8** V 389.]

**Phenosafranine** (3,7-diamino-5-phenylphenazinium chloride) [81-93-6] *M* **322.8**, *m* **>300°**,  $\lambda_{\text{max}}$  **530nm ( $\text{H}_2\text{O}$ )**. Crystallise the chloride from dilute HCl. It has UV with  $\lambda_{\text{max}}$  at 530nm in  $\text{H}_2\text{O}$ . The *picrate* decomposes on heating and has a solubility of 0.0048% in  $\text{H}_2\text{O}$  at  $18^\circ$ . [Beilstein **23** H 395, *Beilstein* **25** H 394, **25** I 654, **25** II 338, **25** III/IV 3050.]

**Phenothiazine** [92-84-2] **M 199.3, m 184-185°**. Crystallise it from \*benzene, toluene, hexane or Me<sub>2</sub>CO (charcoal) after boiling for 10 minutes under reflux. Filter the crystals off and dry them in an oven at 100°, then in a vacuum desiccator over paraffin chips. Also recrystallise it twice from water and dry it in an oven at 100° for 8-10 hours. It sublimes at 130°/1mm and has UV with  $\lambda_{\text{max}}$  at 253nm in heptane. [Beilstein 27 I 225, 27 II 32, 27 III/IV 1214.]

**Phenoxazine** [135-67-1] **M 199.2, m 156°, 156-158°, 158-159°, b 215°/4mm**. Crystallise phenoxazine from EtOH and sublime it *in vacuo*. If too impure then extract it in a Soxhlet extractor using toluene. Evaporate the solvent and dissolve the residue (*ca* 100g) in \*C<sub>6</sub>H<sub>6</sub> (1L), **CARCINOGEN** (use an efficient fume cupboard) and chromatograph it through an Al<sub>2</sub>O<sub>3</sub> column (50 x 450 mm) using \*C<sub>6</sub>H<sub>6</sub>. The eluent (*ca* 3L) is evaporated to *ca* 150ml and cooled when *ca* 103g of phenoxazine **m 149-153°** are obtained. Sublimation yields platelets **m 158-159°**. It forms a green *picrate* **m 141.5-142°**. [Gilman & Moore *J Am Chem Soc* 79 3485 1957, Müller et al. *J Org Chem* 24 37 1959, Beilstein 27 I 223, 27 III/IV 1209.]

**2-Phenyl-1-azaindolizine** [4105-21-9] **M 194.2, m 140°, pK<sub>Est</sub> ~1.9**. Crystallise the indolizine from EtOH, \*benzene/petroleum ether, hexane (**m 135-136°**) or cyclohexane (**m 136-137°**). The *hydrochloride 2H<sub>2</sub>O* has **m 114-116°** (from H<sub>2</sub>O) and the *picrate* has **m 228-229°** (from AcOH) and 236-238° (from Me<sub>2</sub>CO). [Adams & Dix *J Am Chem Soc* 80 4618 1958, Beilstein 23 III/IV 1705.]

**Phenylbutazone (4-butyl-1,2-diphenylpyrazolidin-3,5-dione)** [50-33-9] **M 308.4, m 105°, 106-108°**. Crystallise the dione from EtOH. Its pK<sup>23</sup> is 4.52 (in H<sub>2</sub>O), 4.89 (in 50% aqueous EtOH) and 5.25 (80% 2-methoxyethanol). It complexes with Hg<sup>2+</sup>, Cd<sup>2+</sup> and Zn<sup>2+</sup>. It has UV with  $\lambda_{\text{max}}$  at 239.5nm in MeOH+50% aqueous HClO<sub>4</sub> and 264nm in aqueous 0.1N NaOH. [Beilstein 24 III/IV 1123.]

**2-Phenyl-1,3-diazahexahydroazulene** [2161-31-1] **M 212.3**. Recrystallise the azulene three times from de-aerated cyclohexane in the dark.

**2-Phenylindolizine** [25379-20-8] **M 193.2, m 214°(dec), pK<sub>Est</sub> ~4.4**. Crystallise 2-phenylindolizine from EtOH. The *0.25HCl* crystallises from MeCN with **m 109°**, and the *picrate* has **m 161°** when crystallised from EtOAc. [Beilstein 20 II 304, 20 III/IV 4033, 20/8 V 244.]

**1-Phenyl-5-mercaptotetrazole** [86-93-1] **M 178.2, m 150° (dec), 155° (dec), 157-158°, pK<sup>25</sup> 3.65 (5% aqueous EtOH)**. Purify the tetrazole by recrystallisation from EtOH or CHCl<sub>3</sub> (**m 152°**) [Tautomerism: Kauer & Sheppard *J Org Chem* 32 3580 1967, UV: Leiber et al. *Can J Chem* 37 563 1959]. The *ammonium salt* crystallises from EtOH and decomposes at 176°. The *sodium salt* crystallises from EtOH/\*C<sub>6</sub>H<sub>6</sub>, melts at 96° and decomposes at 145° [Stollé *J Prakt Chem* [2] 133 60 1932]. It is used for the determination of Bi and Pd. [Fresenius *Z Anal Chem* 261 151 1972, Beilstein 26 III/IV 2065.]

**4-(3-Phenylpropyl)pyridine [1-phenyl-3(4-pyridyl)propane]** [2057-49-0] **M 197.3, b 150-152°/5mm, 322°/atm, d<sub>4</sub><sup>25</sup> 1.03, n<sub>D</sub><sup>20</sup> 1.563, pK<sub>Est</sub> ~6.0**. This is prepared from the potassium salt of 4-picoline [obtained by dissolving K metal (1 mol) and a few mg of ferric oxide as catalyst in liquid NH<sub>3</sub> until the blue colour is discharged, indicating the formation of KNH<sub>2</sub>, to which is added 4-picoline (1 mol) rapidly forming a reddish-amber solution] and phenethyl bromide (1 mol, see [103-63-9]), allowing the NH<sub>3</sub> to evaporate (6-10 hours, in an efficient fume cupboard), adding H<sub>2</sub>O, the mixture is extracted into Et<sub>2</sub>O, evaporated to dryness and the residual oil is distilled, preferably under a vacuum to give the desired *phenylpropylpyridine*, **b 150-152°/5-6mm**, in 56% yield. The *hydrochloride*, obtained by evaporating a concentrated HCl solution of the base on a steam bath and the residual solid, recrystallised from EtOH/Et<sub>2</sub>O, has **m 143.5°** and is hygroscopic. [Bergstrom et al. *J Org Chem* 10 452 1945, Beilstein 20 III/IV 3687.]

**4-Phenylpyridine-2-carbonitrile** [18714-16-4] **M 180.1, m 97-101°, pK<sub>Est</sub> <0**. Purify the nitrile by recrystallisation from petroleum ether (**m 99-100°**). [Case & Gasper *J Am Chem Soc* 78 5842 1956, Beilstein 22 III/IV 1261.]

**Phenyl 2-pyridyl ketoxime** [1826-28-4] M 198.2, m 151-152°, 154-156°, pK<sub>1</sub><sup>25</sup> 3.84, pK<sub>2</sub><sup>25</sup> 10.71 for *E*-isomer. The *E*-isomer crystallises from EtOH (charcoal). It isomerises to the *Z*-isomer on melting or in boiling *o*-xylene, and crystallises from EtOH or cyclohexanol with m 166-168°. [Beilstein 21 H 330, 21 III/IV 4120, 21/8 V 568.]

**6-Phenylquinoline** [612-95-3] M 205.3, m 110.5-111.5°, pK<sub>Est</sub> ~5.2. Crystallise 6-phenylquinoline from EtOH (charcoal). The *picrate* has m 105° (from Me<sub>2</sub>CO). It has UV with λ<sub>max</sub> at 253nm in aqueous EtOH and 263 and 325nm in aqueous HCl. [Beilstein 20 H 483, 20 III/IV 4151.]

**2-Phenylquinoline-4-carboxylic acid (Cinchophen, Atophen)** [132-60-5] M 249.3, m 215°, pK<sub>Est(1)</sub> ~0.5 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~5.1 (N). It crystallises from EtOH (ca 20ml/g), in several modifications with m 196° and 216°(subliming at 65°). [Beilstein 22 H 103, 22 II 518, 22 II 70, 22 III/IV 1358.]

**1-Phenyl-5-sulfanilamidopyrazole** [526-08-9] M 314.3, m 177-178°, 178-179°. Crystallise it from EtOH or aqueous EtOH. [Schmidt & Druey *Helv Chim Acta* 41 309 1958, Beilstein 25 III/IV 2029.]

**4-Phenyl-1,2,4-triazole-3,5-diol (4-phenylurazole)** [15988-11-1] M 175.2, m 207-209°. Crystallise 4-phenylurazole from water or 95% EtOH. Dissolve 35g in 80ml of boiling 95% EtOH and on cooling 90-95% is recovered with m 209-210°. It has IR with ν<sub>max</sub> at 1685 and 3120cm<sup>-1</sup>. [Cookson et al. *Org Synth* 51 121 1971, Beilstein 26 I 64, 26 III/IV 540.]

**4-Phenyl-1,2,4-triazoline-3,5-dione (PTAD)** [4233-33-4] M 175.2, m 165-170°(dec), 170-177°(dec). PTAD forms carmine red needles by sublimation (ice cold finger) at 100°/0.1mm, and/or by recrystallisation from EtOH. Its IR has ν<sub>max</sub> at 1760 and 1780 cm<sup>-1</sup>. [Cookson et al. *Org Synth* 51 121 1971, Moore et al. *J Org Chem* 39 3700 1974, Beilstein 26 I 57, 26 III/IV 540.]

**9-Phenyl-9-xanthenol (hydroxypixyl)** [596-38-3] M 274.3, m 158-161°, 158.5-159°, 159°. Dissolve hydroxypixyl in AcOH and add H<sub>2</sub>O whereby it separates as colourless prisms. It is slightly soluble in CHCl<sub>3</sub>, soluble in \*C<sub>6</sub>H<sub>6</sub> but insoluble in petroleum ether. It sublimes on heating. Its UV in H<sub>2</sub>SO<sub>4</sub> has λ<sub>max</sub> at 450nm (ε 5620) and 370nm (ε 24,900) and the HClO<sub>4</sub> salt in CHCl<sub>3</sub> has λ<sub>max</sub> at 450 (ε 404) and 375nm (ε 2420). [Sharp *J Chem Soc* 2558 1958, Bünzly & Decker *Chem Ber* 37 2983 1904, Chattopadhyaya & Reece *J Chem Soc, Chem Commun* 639 1978, Gomberg & Cone *Justus Liebigs Ann Chem* 370 142 1909, Beilstein 17 I 80, 17 II 161, 17 III/IV 1704, 17/4 V 675.]

**“Phosphine” [dye CI 793, Chrysaniline mononitrate, 3-amino-9-(4-aminophenyl)-acridinium mononitrate]** [10181-37-0] M 348.4, m >250°(dec), pK<sup>20</sup> 7.71 (50% aqueous EtOH). Crystallise the dye from \*benzene/EtOH. The *free base* crystallises from \*C<sub>6</sub>H<sub>6</sub> in yellow crystals m 229-230°. [Albert & Goldacre *J Chem Soc* 709 1946, Beilstein 22 H 91, 22 I 651, 22 II 403, 22 III/IV 5513.]

**Phthalazine** [253-52-1] M 130.2, m 90-91°, b 190°/30mm, 175°/17mm, pK<sup>20</sup> 3.47. Phthalazine crystallises from diethyl ether or \*benzene, and sublimes under a vacuum. The *hydrochloride* forms needles from EtOH with m 235-236°(dec) and the *picrate* has m 208-210°. [Armarego *J Appl Chem* 11 70 1961, Beilstein 23 H 174, 23 III/IV 1233.]

**Phthalazine-1,4-dione (phthalhydrazide)** [1445-69-8] M 162.2, m 330-333°, 336°, 346°, pK<sub>1</sub><sup>20</sup> -3.29 pK<sub>2</sub><sup>20</sup> -0.99, pK<sub>3</sub><sup>20</sup> 5.67, pK<sub>4</sub><sup>20</sup> 13.0. Recrystallise it twice from 0.1M KOH [Merenyi et al. *J Am Chem Soc* 108 7716 1986], EtOH or dimethylformamide and it sublimes >300°. [Beilstein 24 H 371, 24 II 194.]

**Phthalazone (1-hydroxyphthalazine)** [119-39-1] M 146.2, m 183-184°, 186-188°, b 337°/760mm, pK<sub>1</sub><sup>20</sup> -2.2, pK<sub>2</sub><sup>20</sup> -1.4, pK<sub>3</sub><sup>20</sup> 11.99. Phthalazone crystallises from H<sub>2</sub>O or EtOH and sublimes *in vacuo*. [Beilstein 24 H 142, 24 III/IV 400.]

**2-Picoline-*N*-oxide (2-methylpyridine-1-oxide)** [931-19-1] M 109.1, m 41-45°, b 89-90°/0.8-0.9mm, 90-100°/1mm, 110°/4mm, 135°/5mm, 123°/9mm, 123-124°/15mm, 259-261°/atm, n<sub>D</sub><sup>25</sup> 1.5854 (supercooled),

**pK<sup>25</sup> 1.10.** Purify the *N*-oxide by fractional distillation, and it can be recrystallised from \*C<sub>6</sub>H<sub>6</sub>/hexane but is *hygroscopic*. [Bullitt & Maynard *J Am Chem Soc* **76** 1370 1954, Ross et al. *J Am Chem Soc* **78** 3625 1956, IR: Wiley & Slaym Aker *J Am Chem Soc* **79** 2233 1957.] The *picrate* has **m** 125-126.5° (from EtOH) [Boekelheide & Linn *J Am Chem Soc* **76** 1286 1954]. The *phthalate* has **m** 115-116° (from EtOH) [den Hertog et al. *Rec Trav Chim Pays Bas* **70** 591 1951]. [Beilstein **20** III/IV 2689, **20/5** V 479.]

**3-Picoline-*N*-oxide (3-methylpyridine-1-oxide)** [1003-73-2] **M 109.1, m 37-39°, 37-38° (evacuated capillary), 84-85°/0.3mm, 101-103°/0.7-0.8mm, 114-115°/1.5mm, 118°/2mm, pK<sup>25</sup> 1.08.** Purify the *N*-oxide by careful fractionation *in vacuo*. The distillate remains supercooled for several days before solidifying. It is a slightly *hygroscopic* solid which could melt in the hand. The *picrate* has **m** 149-151° (from EtOH). [Taylor & Corvetti *Org Synth Coll Vol IV* 654 1963, IR: Katritzky et al. *J Chem Soc* 3680 1959, Jaffé & Doak *J Am Chem Soc* **77** 4441, 4481 1955, Boekelheide & Linn *J Am Chem Soc* **76** 1286 1954]. [Beilstein **20** III/IV 2719, **20/5** V 517.]

**4-Picoline-*N*-oxide (4-methylpyridine-1-oxide)** [1003-67-4] **M 109.1, m 182-184°, 185-186°, 186-188°, pK<sup>25</sup> 1.29.** Recrystallise the *N*-oxide from EtOH/EtOAc, Me<sub>2</sub>CO/Et<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>. [Bullitt & Maynard *J Am Chem Soc* **76** 1370 1954, Boekelheide & Linn *J Am Chem Soc* **76** 1286 1954]. [Beilstein **20** III/IV 2741, **20/5** V 558.]

**Picolinic acid (pyridine-2-carboxylic acid)** [98-98-6] **M 123.1, m 138°, pK<sub>1</sub><sup>25</sup> 1.03 (1.36), pK<sub>2</sub><sup>25</sup> 5.30 (5.80).** Crystallise the acid from water or \*benzene. The *picrate* has **m** 185-187° (from MeOH). [Beilstein **22** H 33, **22** I 502, **22** II 30, **22** III/IV 303, **22/2** V 3.]

**α-Picolinium chloride** [14401-91-3] **M 129.6, m 77-78°, 85°, b 229°/atm.** A 1:1 mixture of α-picoline and HCl is distilled at 275°, then sublimed in a vacuum at 91-91.5°. [Beilstein **20** H 236, **20** III/IV 2685, **20/5** V 474.]

***N*-4-Picolinoylbenzimidazole** [100312-29-6] **M 173.3, m 105-107°.** Recrystallise the imidazole three times from hexane [Fife & Przystas *J Am Chem Soc* **108** 4631 1986].

**Picrolic acid [3-methyl-4-nitro-1-(4-nitrophenyl)-2-pyrazolin-5-one, picrolonic acid]** [550-74-3] **M 264.2, m 120°(dec), 116.5°(dec at 125°) 125°.** Crystallise picrolic acid from water or EtOH (solubility is 0.123% at 15° and 1.203% at 100° in H<sub>2</sub>O; and 1.107% at 0° and 11.68% at 81° in EtOH). It forms Ca, Cu Hg, Mg, Na, Sr, Pb and many other metal complexes [Maquestian et al *Bull Soc Chim Belg* **82** 233 1973, Isaki et al. *Chem Ber* **74** 1420 1941]. [Beilstein **24** H 51, **24** I 218, **24** II 25, **22** III/IV 105.]

**Pinacyanol chloride (Quinaldine Blue)** [2768-90-3] **M 388.9, CI 808, m 270°(dec).** Crystallise the dye from EtOH/diethyl ether. The *iodide* [605-91-4] crystallises from MeOH or EtOH with **m** 298-299°(dec). [Beilstein **23** H 320, **23** I 89, II 282, III/IV 2064, **23/10** V 129.]

***dl*-Pipicolinic acid (piperidine-2-carboxylic acid)** [4043-87-2] **M 129.1, m 264°, 280°(dec), pK<sub>1</sub><sup>25</sup> 2.29, pK<sub>2</sub><sup>25</sup> 10.77.** It crystallises from water. The (±)-*picrate* has **m** 158-159° (from EtOH or \*C<sub>6</sub>H<sub>6</sub>). [Beilstein **22** H 7, **22** III/IV 97, **22/1** V 220.] The *R*(+)-*enantiomer* [1723-00-8] has **m** 277°(dec) and [α]<sub>D</sub><sup>20</sup> +27° (c 4, H<sub>2</sub>O), and the *S*(-)-*enantiomer* [3105-95-1] has **m** 277°(dec) and [α]<sub>D</sub><sup>20</sup> -26° (c 4, H<sub>2</sub>O). [cf p 603, Beilstein **22** III/IV 96, **22/1** V 220.]

**Piperazine** [110-85-0] **M 86.1, m 110-112°, 44° (hexahydrate 142-63-2) b 125-130°/760mm, pK<sub>1</sub><sup>25</sup> 5.33, pK<sub>2</sub><sup>25</sup> 9.73.** Piperazine crystallises from EtOH or anhydrous \*benzene and is dried at 0.01mm. It can be sublimed under vacuum and purified by zone melting. The *hydrochloride* has **m** 172-174° (from EtOH), and the *dihydrochloride* crystallises from aqueous EtOH and has **m** 318-320° (dec, sublimes at 295-315°). The *picrate* has **m** ~200°, and the *picrolonate* crystallises from dimethylformamide (**m** 259-261°). [Beilstein **23** H 4, **23** I 4, **23** II 3, **23** III/IV 15, **23/1** V 30.]

§ Piperazine on polystyrene support is commercially available.

**Piperazine-*N,N'*-bis(2-ethanesulfonic acid) (PIPES)** [5625-37-6] **M 302.4**,  $\text{pK}_1^{25} < 3$ ,  $\text{pK}_2^{25} 6.82$  (7.82). Purify PIPES from boiling water (maximum solubility is about 1g/L) or as described for ADA (see [26239-55-4] in "Aliphatic Compounds" in this Chapter). [Good et al. *Biochemistry* 5 469 1966, *Beilstein* 23/12 V 380.]

**Piperazine dihydrochloride (H<sub>2</sub>O)** [142-64-3 (2HCl); 6094-40-2 (xHCl)] **M 177.1**, **m 82.5-83.5°**. Crystallise the salt from aqueous EtOH and dry it at 110°. [*Beilstein* 23 III/IV 17, 23/1 V 30.]

**Piperazine phosphate (H<sub>2</sub>O)** [18534-18-4] **M 197.6**. Crystallise it twice from water, air-dry and store for several days over Drierite. The salt dehydrates slowly if heated at 70°. [*Beilstein* 23 III/IV 18, 23/1 V 30.]

**Piperidine** [110-89-4] **M 85.2**, **f -9°**, **b 35.4°/40mm**, **106°/760mm**,  $d_4^{20} 0.862$ ,  $n_D^{20} 1.4535$ ,  $n_D^{25} 1.4500$ ,  $\text{pK}^{25} 11.20$  (**basic**). Dry piperidine with BaO, KOH, CaH<sub>2</sub>, or sodium, and fractionally distil (optionally from sodium, CaH<sub>2</sub>, or P<sub>2</sub>O<sub>5</sub>). Purify from pyridine by zone melting. [*Beilstein* 22 H 6, 22 I 5, 22 II 3, 22 III/IV 287, 22/2 V 3.]

§ Piperidine on polystyrene support is commercially available.

**Piperidine-2,6-dione (glutarimide)** [1121-89-7] **M 113.1**, **m 163-165°**, **155-157°**, **154°**, **152-154°**,  $\text{pK}^{25} 11.43$  (**acidic**). Purify it by dissolving 75g in 200ml of H<sub>2</sub>O, boil for 30 minutes with 2g of charcoal, filter, evaporate to dryness and recrystallise the residue from 125m L of 95% EtOH to give 70g of white crystals, **m 152-154°**. It also crystallises from Me<sub>2</sub>CO (**m 163-165°**) or EtOH (**m 153-154°**). The *N*-bromo derivative (a brominating agent) crystallises from H<sub>2</sub>O with **m 180-185°**. [Paris et al. *Org Synth Coll Vol IV* 496 1963, *Beilstein* 21 H 382, 21 I 331, 21 II 307, 21 III/IV 4582.]

**Piperidinium chloride** [6091-44-7] **M 121.6**, **m 244-245**, **247-248°**. Crystallise the salt from EtOH/diethyl ether in the presence of a small amount of HCl. [*Beilstein* 20 H 6, 20 II 10, 20 III/IV 295, 20/2 V 13.]

**Piperidinium nitrate** [6091-45-8] **M 145.2**, **m 141°**, (**155-157° from EtOH**). The nitrate crystallises from acetone/ethyl acetate or EtOH. [*Beilstein* 20 H 12, 20 II 10, 20 III/IV 295, 20/2 V 14.]

**4-Piperidone [piperidin-4(1H)-one,  $\gamma$ -]** [41661-47-6] **M 99.1** (**anhydrous**), **cannot be distilled**,  $\text{pK}^{25} 8.6$  (**acidic**). It is a yellow irritating oil which is an alkaloid in the leaves and branches of *Dichilus* species (leguminosae), e.g. *D. strictus*, *D. gracilis*, *D. lebeckioides*, *D. pilosus* and *D. reflexus*. It decomposes on distillation and is purified and stored as the *monohydrate monohydrochloride* [*anhydrous* 41979-39-9; *hydrate* 40064-34-4], **M 135.6** (*anhydrous salt*), which is in fact *4,4-dihydroxy-piperidinium chloride* that recrystallises from H<sub>2</sub>O with **m 94-96°**, **97-100°** (+ 1 H<sub>2</sub>O), or from EtOH/Et<sub>2</sub>O with **m 139-141°** (+ 1.5 EtOH); and the *anhydrous salt* has **m 147-149°**. See above for the *N*-acetyl, *N*-benzyl and *N*-methyl derivatives. The *N*-benzoyl derivative [24686-78-0], **M 203.2**, has **m 49-50°** and **b 158-160°/0.2mm**, and the *oxime* [79858-41-6], **M 114.1**, crystallises from dry \*C<sub>6</sub>H<sub>6</sub> in needles **m 117-118°** (anhydrous), and is hygroscopic. [*Beilstein* 21/6 V 419.]

**Piperine (1-piperoylpiperidine)** [94-62-2] **M 285.4**, **m 129-129.5°**,  $\text{pK}^{15} 1.98$ . Piperine crystallises as light yellow crystals from EtOH or EtOAc (**m 132°**), aqueous EtOH (**m 128-129°**), Et<sub>2</sub>O (**m 129°**), or \*benzene/ligroin. [*Beilstein* 20 H 79, 20 I 23, 20 II 53, 20 III/IV 1341, 20/3 V 469.]

**Poly(*N*-vinylcarbazole)** [25067-59-8]. Precipitate it seven times from tetrahydrofuran with MeOH, with final freeze-drying from \*benzene. Dry it under vacuum.

**Poly(4-vinylpyridine)** [25232-41-1] **M (105.1)<sub>n</sub>**. Purify it by repeated precipitation from solutions in EtOH and dioxane, and then EtOH and ethyl acetate. Finally, freeze-dry a *tert*-butanol solution.

**Poly(*N*-vinylpyrrolidone)** [9003-39-8] **M (111.1)<sub>n</sub>**, **crosslinked** [25249-54-1] **m >300°**. Purify it by dialysis, and freeze-drying. Also by precipitation from CHCl<sub>3</sub> solution by pouring into ether. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. For the crosslinked polymer, purification is by boiling for 10 minutes in 10% HCl and then washing with glass-distilled water until free from Cl ions. Finally, Cl ions are removed more readily by neutralising with

KOH and continued washing.

**(±)-Primaquine diphosphate** (*RS*- 8-[4-amino-1-methylbutylamino]-6-methoxyquinoline di-phosphate) [63-45-6] **M** 455.4, **m** 197-198°, 204-206°(dec),  $\text{pK}_{\text{Est}(1)} \sim 3.38$  (ring  $\text{N}^+$ ),  $\text{pK}_{\text{Est}(2)} \sim 10.8$  ( $\text{NH}_3^+$ ). It forms yellow crystals from 90% aqueous EtOH and is moderately soluble in  $\text{H}_2\text{O}$ . The *oxalate salt* has **m** 182.5-185° (from 80% aqueous EtOH), and the *free base* is a viscous liquid **b** 165-170°/0.002mm, 175-177°/2mm. [Elderfield et al. *J Am Chem Soc* **68** 1526 1964, Elderfield et al. *J Am Chem Soc* **77** 4817 1955, *Beilstein* **22** III/IV 5817.]

**Proclavine (3,6-diaminoacridine)** [92-62-6] **M** 209.2, **m** 284-286°,  $\text{pK}_1^{25} -2.7$ ,  $\text{pK}_2^{25} 0.55$ ,  $\text{pK}_3^{25} 9.49$ . It crystallises from aqueous MeOH. The *picrate* crystallises from aqueous pyridine with **m**  $\sim 185^\circ$ . [*Beilstein* **22** H 487, **22** I 649, **22** II 397, **22** III/IV 5487, **22/11** V 322.] For proflavin see 3,6-diaminoacridine hydrochloride.

**Propidium iodide (3,8-diamino-5-(3-diethylaminopropyl)-6-phenylphenanthridinium iodide methiodide)** [25535-16-4] **M** 668.4, **m** 210-230°(dec),  $\text{pK}_{\text{Est}(1)} \sim 4$  (aniline  $\text{NH}_2$ ),  $\text{pK}_{\text{Est}(2)} \sim 8.5$  ( $\text{EtN}_2$ ). It crystallises as red crystals from  $\text{H}_2\text{O}$  containing a little KI. It fluoresces strongly with nucleic acids. [Watkins *J Chem Soc* 3064 1952, *Beilstein* **22** III/IV 5519.] **TOXIC.**

**(±)-Propylene carbonate (4-methyl-1,3-dioxolan-2-one)** [108-32-7] **M** 102.1, **b** 79-80°/0.08mm, 110°/0.5-1mm, 112-114°/2mm, 241°/760mm,  $d_4^{20} 1.204$ ,  $n_D^{20} 1.423$ . It is manufactured by reaction of 1,2-propylene oxide with  $\text{CO}_2$  in the presence of a catalyst (quaternary ammonium halide). Contaminants include propylene oxide, carbon dioxide, 1,2- and 1,3-propanediols, allyl alcohol and ethylene carbonate. It can be purified by percolation through molecular sieves (Linde 5A, dried at 350° for 14 hours under a stream of argon), followed by distillation under a vacuum. [Jasinski & Kirkland *Anal Chem* **39** 163 1967.] It can be stored over molecular sieves under an inert gas atmosphere. When purified in this way it contains less than 2 ppm of water. Activated alumina and dried CaO have also been used as drying agents prior to fractional distillation under reduced pressure. It has been dried with 3A molecular sieves and distilled under nitrogen in the presence of *p*-toluenesulfonic acid, then redistilled and the middle fraction collected. [*Beilstein* **19** III/IV 1564, **19/4** V 21.]

**dl-Propylene oxide** [75-56-9] **M** 58.1, **b** 34.5°,  $d_4^{20} 0.829$ ,  $n_D^{20} 1.3664$ . Dry the oxide with  $\text{Na}_2\text{SO}_4$  or  $\text{CaH}_2$  and fractionally distil it through a packed column (glass helices), after refluxing with Na,  $\text{CaH}_2$ , or KOH pellets. [*Beilstein* **17** I 4, **17** II 131, **17** III/IV 17, **17/1** V 17.] The *R*(+)-enantiomer [15448-47-2] and the *S*(-)-enantiomer [16088-62-3] have **b** 33-34°/atm and  $[\alpha]_D^{20}$  (+)14.6° and (-)14.6° (neat). [*Beilstein* **17/1** V 17.]

**6-Propyl-2-thiouracil (propacil, propylail)** [51-52-5] **M** 170.2, **m** 218-220°, 218-220°,  $\text{pK}_1^{21} -6.54$  (aqueous  $\text{H}_2\text{SO}_4$ ),  $\text{pK}_2^{21} -4.22$  (aqueous  $\text{H}_2\text{SO}_4$ ),  $\text{pK}_3^{21} 8.25$  (4% aqueous EtOH). Purify propacil by recrystallisation from  $\text{H}_2\text{O}$  (soluble in 900 parts at 20°, and 100 parts at 100°). Its UV (MeOH) has  $\lambda_{\text{max}}$  at 277nm. [Anderson et al. *J Am Chem Soc* **67** 2197 1945, Vanderhaegue *Bull Soc Chim Belg* **59** 689 1950, *Beilstein* **24** III/IV 1333.]

**Protopine [fumarine, macleyine, 4,6,7,14-tetrahydro-5-methyl-bis[1,3]-benzodioxolo[4,5-*c*:5',6'-*g*]azecine-13(5*H*)-one]** [130-86-9] **M** 353.4, **m** 208°, 209°, 211°,  $\text{pK}^{25} 5.99$ . It crystallises from EtOH/ $\text{CHCl}_3$ . The *picrate* has **m**  $\sim 240^\circ$ (dec). [*Beilstein* **27** H 558, **17** I 568, **17** II 620, **17** III/IV 6881.]

**Pteridine** [91-18-9] **M** 132.2, **m** 139.5-140°,  $\text{pK}_1^{20} 4.05$  (equilibrium, hydrate),  $\text{pK}_2^{20} 11.90$  (OH of hydrate). It crystallises from EtOH, \*benzene, *n*-hexane, *n*-heptane or petroleum ether. It is best purified by sublimation at 120-130°/20mm. Store at 0°, in the dark. The yellow crystalline solid turns green in the presence of light and on long standing in the dark, and sublimation leaves the dark-coloured material behind. [Albert et al. *J Chem Soc* 474 1951; see Albert & Armarego *Adv Heterocycl Chem* **4** 1 1965, *Beilstein* **26** III/IV 1770.]

**2,4-(1*H*,3*H*)-Pteridinedione  $\text{H}_2\text{O}$  (lumazine)** [487-21-8] **M** 182.1, **m**  $>350^\circ$ ,  $\text{pK}_1^{20} <1.0$ ,  $\text{pK}_2^{20} 7.94$ . Crystallise the dione from water. It has also been purified as for pterin [2236-60-4] below. [Dallacker & Steiner *Justus Liebigs Ann Chem* **660** 98 1962, *Beilstein* **26** III/IV 2489.]

**Pterin (2-aminopteridin-4(3H)-one)** [2236-60-4] **M 163.1, m >300°**,  $\text{pK}_1^{20}$  2.27 (basic),  $\text{pK}_2^{20}$  7.96 (acidic). It is dissolved in hot 1% aqueous ammonia, filtered, and an equal volume of hot 1M aqueous formic acid is added. The solution is allowed to cool at 0-2° overnight. The solid is collected and washed with distilled water several times by centrifugation and dried *in vacuo* over P<sub>2</sub>O<sub>5</sub> overnight, and then at 100° overnight (any ammonium formate in the sample evaporates off). [Beilstein 26 III/IV 3936.]

**Pterocarpin** {(6a*R*-*cis*)-6a,12a-dihydro-3-methoxy-6*H*-[1,3]dioxolo[5,6]benzofuro[3,2*c*][1]-benzopyran} [524-97-0] **M 298.3, m 165°, 165-166°,  $[\alpha]_{546}^{20}$  -215° (c 0.5, CHCl<sub>3</sub>)**. Crystallise it from EtOH, or petroleum ether. [Fukui & Nakayama *Bull Chem Soc Jpn* 42 1408 1969, Packler & Underwood *Tetrahedron* 23 1817 1967, Beilstein 19 II 459, 19 III/IV 5789.]

**Pterioic acid (2-amino-6-*p*-carboxyanilinomethylpteridin-4(3H)-one)** [119-24-4] **M 312.3, m >300°(dec)**,  $\text{pK}_{\text{Est}(1)} \sim 2.3$  (basic, N1),  $\text{pK}_{\text{Est}(2)} \sim 2.6$  (basic, CH<sub>2</sub>NH),  $\text{pK}_{\text{Est}(3)} \sim 4.5$  (COOH),  $\text{pK}_{\text{Est}(4)} \sim 7.9$  (acidic 4-OH). Crystallise it from dilute HCl. Dry it *in vacuo*. Hygroscopic **IRRITANT**. [Nair et al. *J Org Chem* 46 3152 1981, Beilstein 26 III/IV 3942.]

**Purine** [120-73-0] **M 120.1, m 216-217°,  $\text{pK}_1^{20}$  2.30,  $\text{pK}_2^{20}$  9.86**. It crystallises from toluene or EtOH, and sublimes at 100-150°/0.1mm or 160°/10<sup>-4</sup>mm. The *picrate* has **m 207-209°** after crystallisation from 20 volumes of H<sub>2</sub>O. [Beilstein 26 H 354, 26 III/IV 1736.]

**Pyocyanine (1-hydroxy-5-methylphenazinium zwitterion)** [85-66-5] **M 210.2, m 133° (sublimes and decomposes on further heating),  $\text{pK}^{25} -3.5$** . It crystallises from H<sub>2</sub>O as dark blue needles. The *picrate* has **m 190° (dec)**. [Beilstein 23 H 395, 23 I 59, 23 II 234, 23/8 V 395.]

**Pyrazine** [290-37-9] **M 80.1, m 47°, 57°, b 115.5-115.8°,  $\text{pK}_1^{20}$  -6.25 (aqueous H<sub>2</sub>SO<sub>4</sub>),  $\text{pK}_2^{20}$  1.1 (0.51 at 20°)**. Distil pyrazine in steam and crystallise it from water. Purify also by zone melting. [Beilstein 23 H 91, 23 II 80, 23 III/IV 899, 23/5 V 351.]

**Pyrazinecarboxamide** [98-96-4] **M 123.1, m 189-191° (sublimes slowly at 159°),  $\text{pK}^{25} -0.5$** . The amide crystallises from water, EtOH or 1:1 hexane/EtOH in four modifications *viz* α-form, β-form, δ-form and γ-form. [Rø & Sørum *Acta Cryst* 28B 1677 1972, Beilstein 25 III/IV 772.]

**Pyrazinecarboxylic acid** [98-97-5] **M 124.1, m 225-229°(dec),  $\text{pK}_1^{25} -3.0$ ,  $\text{pK}_2^{25} -0.7$ ,  $\text{pK}_3^{25} 2.70$** . It crystallises from water. The *methyl ester* has **m 62°** (from petroleum ether). [Sauville & Spoerri *J Am Chem Soc* 63 3153 1941, Beilstein 25 III/IV 771.]

**Pyrazine-2,3-dicarboxylic acid** [89-01-0] **M 168.1, m 183-185°(dec), 187°(dec),  $\text{pK}_1 < -2.0$ ,  $\text{pK}_2$  0.9,  $\text{pK}_3$  2.77 (2.20)**. Crystallise the dicarboxylic acid from water and dry it at 100°. The *dimethyl ester* has **m 62-63°** (from Et<sub>2</sub>O or Et<sub>2</sub>O/petroleum ether). [Beilstein 25 H 168, 25 II 164, 25 III/IV 1064.]

**Pyrazine-1,4-dioxide** [2423-65-6] **M 112.1, m 285-286°,  $\text{pK}_{\text{Est}} < 0$** . If the sample contains pyrazine-1-oxide, then place it in a Soxhlet extractor and extract it with hot petroleum ether (b 60-68°) in which the mono-oxide is soluble. Collect the dioxide from the thimble and recrystallise it from MeOH. It is dried *in vacuo*. It has  $\nu_{\text{max}}$  at 1270cm<sup>-1</sup>. [Koelsch & Gumprecht *J Org Chem* 23 1605 1958.]

**Pyrazine-1-oxide** [2433-84-9] **M 96.1, m 113-114°,  $\text{pK}_{\text{Est}} < 0$** . Recrystallise the oxide from \*C<sub>6</sub>H<sub>6</sub>. It is soluble in hot petroleum ether (b 60-68°, see pyrazine-1,4-dioxide above). Dry it *in vacuo*. It has  $\nu_{\text{max}}$  at 1305cm<sup>-1</sup>. [Koelsch & Gumprecht *J Org Chem* 23 1605 1958.]

**Pyrazole** [288-13-1] **M 68.1, m 70°, b 96°/16mm,  $\text{pK}^{25} 2.48$** . Crystallise pyrazole from petroleum ether, cyclohexane, or water. Its solubility in H<sub>2</sub>O at 9.6° is 2.7moles/L, and at 24.8° it is 19.4moles/L; in cyclohexane at 31.8° it is 0.577moles/L, and at 56.2° it is 5.86moles/L; and in benzene at 5.2° it is 0.31moles/L, and at 46.5° it is 16.8moles/1000ml. [Barszcz et al. *J Chem Soc, Dalton Trans* 2025 1986, Beilstein 23 H 39, 23 I 15, 23 II 33, 23 III/IV 550, 23/4 V 122.]

**Pyrazole-*N*-1-carboximidine hydrochloride (Praxadine, 1-amidinopyrazole hydrochloride, 1*H*-pyrazole-1-carboximidine hydrochloride)** [4023-02-3] **M 146.6, m 165-166°**, 167-168°, 167-168.5°, 167-170°, **pK<sub>Est</sub> ~8.5**. The white crystalline hydrochloride, which can be prepared from pyrazole and cyanamide in *p*-dioxane containing HCl, is purified by recrystallisation from dioxane/H<sub>2</sub>O. It is a good guanylation agent for primary and secondary amines [Bernatowicz et al. *J Org Chem* **57** 2497 1992, Bernatowicz et al. *Tetrahedron Lett* **34** 3389 1993]. The *free base* can be obtained by suspending the hydrochloride in CHCl<sub>3</sub> and bubbling NH<sub>3</sub> gas through, whereby NH<sub>4</sub>Cl separates and is filtered off. CHCl<sub>3</sub> is then distilled off and the residue is washed with a little EtOH and dried *in vacuo*. It crystallises from \*C<sub>6</sub>H<sub>6</sub> and has **m 93-96° (94-98.5° and 97-101°** have also been reported). The *picrate* has **m 202-203° (208°** has also been reported). [Bredereck et al. *Chem Ber* **98** 3178 1965, Jones et al. *J Org Chem* **19** 1428 1954, 2497 1992, *Beilstein* **23/4** V131.]

**1*H*-Pyrazole-3-carboxylic acid** [1621-91-6] **M 112.1, m 208-210°, 210-214° (decarboxylates)**, **pK<sub>25</sub> 3.74, pK<sub>Est(2)</sub> ~12 (acidic NH)**. Purify the acid by precipitation from an alkaline solution with mineral acid, and recrystallise it from H<sub>2</sub>O. On heating, it decarboxylates (to pyrazole *m* 70°) more readily than the 4-carboxylic acid below. The *N*-2-*methyl derivative* has **m 222° (from H<sub>2</sub>O)** and **pK<sub>25</sub> 3.27**. [Habraken et al. *Rec Trav Chim Pays Bas* **85** 1194 1966, Jones et al. *J Org Chem* **19** 1428 1954, Allen & Shirley *J Am Chem Soc* **80** 6273 1958, Knorr *Justus Liebigs Ann Chem* **273** 231 1893, Knorr *Justus Liebigs Ann Chem* **279** 217 1894.]

**1*H*-Pyrazole-4-carboxylic acid** [37718-11-9] **M 112.1, m 282° (decarboxylates)**, **pK<sub>Est(1)</sub> ~3.6, pK<sub>Est(2)</sub> ~12 (acidic NH)**. This acid, also obtained by decarboxylating the more water soluble 3,4,5-tricarboxylic acid, crystallises in yellow prisms from H<sub>2</sub>O. [Buchner & Fritsch *Justus Liebigs Ann Chem* **273** 253 1893, Knorr & Rothenburg *Chem Ber* **28** 693 1895, *Beilstein* **25** H 116.]

**Pyrazole-3,5-dicarboxylic acid** [3112-31-0] **M 174.1, m 287-289°(dec), 295-297°(dec)**, **pK<sub>Est(1)</sub> ~1.2 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~3.7 (CO<sub>2</sub>H), pK<sub>Est(3)</sub> ~12 (NH)**. It crystallises from water as a *monohydrate*, or EtOH. Dry it in a vacuum at 100°. The *dimethyl ester* crystallises from \*C<sub>6</sub>H<sub>6</sub> with **m 155°**. [*Beilstein* **25** III/IV 1047.]

**Pyridazine *N*-oxide (pyridazine 1-oxide)** [1457-42-7] **M 96.1, m 38-39°, b 138-140°/4mm, pK<sub>Est</sub> <1**. Purify the oxide by distillation in a vacuum and by sublimation *in vacuo*. When a solution of the oxide in MeOH is treated with an aqueous solution of CuCl<sub>2</sub>, the [C<sub>4</sub>H<sub>4</sub>N<sub>2</sub>O]<sub>2</sub>-CuCl<sub>2</sub>-2H<sub>2</sub>O-*complex* (**m 182-183°**) is formed from which the oxide can be recovered. [Pollak et al. *J Org Chem* **35** 2478 1970, Klinge et al. *Rec Trav Chim Pays Bas* **95** 21 1976, Ohsawa et al. *Tetrahedron Lett* 1979 1978, Koelsch & Gumprecht *J Org Chem* **23** 1605 1958, *Beilstein* **23** III/IV 890.]

**Pyridine** [110-86-1] **M 79.1, f -41.8°, b 115.6°, d<sub>4</sub><sup>20</sup> 0.9831, n<sub>D</sub><sup>20</sup> 1.51021, pK<sub>25</sub> 5.23**. Likely impurities are H<sub>2</sub>O and amines such as the picolines and lutidines. Pyridine is *hygroscopic* and is miscible with H<sub>2</sub>O and organic solvents. It can be dried with solid KOH, NaOH, CaO, or BaO, followed by fractional distillation. Other methods of drying include standing with Linde type 4A molecular sieves, CaH<sub>2</sub> or LiAlH<sub>4</sub>, azeotropic distillation of the H<sub>2</sub>O with toluene or \*benzene, or treated with phenylmagnesium bromide in ether, followed by evaporation of the ether and distillation of the pyridine. A recommended [Lindauer & Mukherjee *Pure Appl Chem* **27** 267 1971] method dries pyridine over solid KOH (20g/Kg) for 2 weeks and fractionally distills the supernatant over Linde type 5A molecular sieves and solid KOH. The product is stored under CO<sub>2</sub>-free nitrogen. Pyridine can be stored in contact with BaO, CaH<sub>2</sub> or molecular sieves. Non-basic materials can be removed by steam distilling a solution containing 1.2 equivalents of 20% H<sub>2</sub>SO<sub>4</sub> or 17% HCl until about 10% of the base has been carried over along with the non-basic impurities. The residue is then made alkaline, and the base is separated, dried with NaOH and fractionally distilled. *The alkali metals, Na, Li or Cs should NOT be used for drying pyridine, and pyridine derivatives, as they form coloured pyridine radical anions leading to bipyridyls.* [see Schmulback et al. *J Am Chem Soc* **90** 6600 1968].

*Alternatively*, pyridine can be treated with oxidising agents. Thus pyridine (800ml) has been stirred for 24 hours with a mixture of ceric sulfate (20g) and anhydrous K<sub>2</sub>CO<sub>3</sub> (15g), then filtered and fractionally distilled. Hurd and Simon [*J Am Chem Soc* **84** 4519 1962] stirred pyridine (135ml), water (2.5L) and KMnO<sub>4</sub> (90g) for 2 hours at 100°, then set aside for 15 hours before filtering off the precipitated manganese oxides. Addition of solid KOH (*ca* 500g) caused pyridine to separate. It was decanted, refluxed with CaO for 3 hours and distilled.

Separation of pyridine from some of its homologues can be achieved by crystallisation of the oxalates. Pyridine is precipitated as its *oxalate* by adding it to the stirred solution of oxalic acid in acetone. The precipitate is filtered, washed with cold acetone, and pyridine is regenerated and isolated. Other methods are based on complex formation with  $\text{ZnCl}_2$  or  $\text{HgCl}_2$ . Heap, Jones and Speakman [*J Am Chem Soc* **43** 1936 1921] added crude pyridine (1L) to a solution of  $\text{ZnCl}_2$  (848g) in 730ml of water, 346ml of conc HCl and 690ml of 95% EtOH. The crystalline precipitate of  $\text{ZnCl}_2 \cdot (\text{pyridine})_2$  was filtered off, recrystallised twice from absolute EtOH, then treated with a conc NaOH solution, using 26.7g of solid NaOH to 100g of the complex. The precipitate was filtered off, and the pyridine was dried with NaOH pellets and distilled. Similarly, Kyte, Jeffery and Vogel [*J Chem Soc* 4454 1960] added pyridine (60ml) in 300ml of 10% (v/v) HCl to a solution of  $\text{HgCl}_2$  (405g) in hot water (2.3l). On cooling, crystals of pyridine- $\text{HgCl}_2$  (1:1) complex separated and were filtered off, crystallised from 1% HCl (to *m* 178.5-179°), washed with a little EtOH and dried at 110°. The free base was liberated by addition of excess aqueous NaOH and separated by steam distillation. The distillate was saturated with solid KOH, and the upper layer was removed, dried further with KOH, then BaO and distilled. Another possible purification step is fractional crystallisation by partial freezing.

Small amounts of pyridine have been purified by vapour-phase chromatography, using a 180-cm column of polyethyleneglycol-400 (Shell 5%) on Embacel at 100°, with argon as carrier gas. The Karl Fischer titration can be used for determining water content. A colour test for pyrrole as a contaminant is described by Biddiscombe et al. [*J Chem Soc* 1957 1954]. The 1:1-hydrochloride crystallises from EtOH with *m* 144°, *b* 218-219°/760mm (see below) and is hygroscopic. The 1:2-hydrochloride has *m* 46° [58888-58-7] and the *picrate* has *m* 165-166° [1152-90-5]. [*Beilstein* **20** H 181, **20** I 54, **20** II 96, **20** III/IV 2205, **20/5** V 160.]

§ Polystyrene-supported pyridine is commercially available.

**Pyridine-2-aldehyde** [1121-60-4] *M* 107.1, *b* 81.5°/25mm, *d*<sub>4</sub><sup>20</sup> 1.121, *n*<sub>D</sub><sup>20</sup> 1.535, *pK*<sub>1</sub><sup>25</sup> 3.84, *pK*<sub>2</sub><sup>25</sup> 12.68. Purification is achieved by bubbling sulfur dioxide into a solution of 50g of the aldehyde in 250ml of boiled water, under nitrogen, at 0°, until precipitation is complete. The bisulfite addition compound is filtered off rapidly and, after washing with a little water, is refluxed in 17% HCl (200ml) under nitrogen until a clear solution is obtained. Neutralisation with  $\text{NaHCO}_3$  and extraction with ether separated the aldehyde which is recovered by drying the extract, then distilling twice, under nitrogen. [Kyte et al. *J Chem Soc* 4454 1960, *Beilstein* **21** I 287, **21** III/IV 3495, **21/7** V 293.]

**Pyridine-3-aldehyde** [500-22-1] *M* 107.1, *b* 89.5°/14mm, *d*<sub>4</sub><sup>20</sup> 1.141, *n*<sub>D</sub><sup>20</sup> 1.549, *pK*<sub>1</sub><sup>20</sup> 3.80, *pK*<sub>2</sub><sup>20</sup> 13.10. Purified as for pyridine-2-aldehyde. [*Beilstein* **21** I 288, **21** III/IV 3517, **21/7** V 334.]

**Pyridine-4-aldehyde** [872-85-5] *M* 107.1, *b* 79.5°/12mm, *d*<sub>4</sub><sup>20</sup> 1.137, *n*<sub>D</sub><sup>20</sup> 1.544, *pK*<sub>1</sub><sup>20</sup> 4.77, *pK*<sub>2</sub><sup>20</sup> 12.20. Purified as for pyridine-2-aldehyde. [*Beilstein* **21** III/IV 2529, **21/7** V 351.]

**Pyridine-2-aldoxime (pyridine-2-carboxaldoxime)** [873-69-8] *M* 122.1, *m* 111-113°, 114°, *pK*<sub>1</sub><sup>25</sup> 3.56, *pK*<sub>2</sub><sup>25</sup> 10.17. Recrystallise it from  $\text{Et}_2\text{O}$ /petroleum ether or  $\text{H}_2\text{O}$ . The *picrate* has *m* 169-171° (from aqueous EtOH). It is used in peptide synthesis. [UV: Grammaticakis *Bull Chem Soc Fr* 109, 116 1956, Ginsberg & Wilson *J Am Chem Soc* **79** 481 1957, Hanania & Irvine *Nature* **183** 40 1959, Green & Saville *J Chem Soc* 3887 1956, *Beilstein E-isomer* **21** I 288, **21** III/IV 3504, **21/7** V 305.]

**Pyridine-3-aldoxime** [1193-92-6] *M* 122.1, *m* 150°, *pK*<sub>1</sub><sup>20</sup> 4.07, *pK*<sub>2</sub><sup>20</sup> 10.39. Crystallise the oxime from water. [*Beilstein E-isomer* **21** III/IV 3521, **21/7** V 339.]

**Pyridine-4-aldoxime** [696-54-8] *M* 122.1, *m* 129°, *pK*<sub>1</sub><sup>20</sup> 4.73, *pK*<sub>2</sub><sup>20</sup> 10.03. Crystallise the oxime from water. [*Beilstein E-isomer* **21** III/IV 3533, **21/7** V 355.]

**2,6-Pyridinedialdoxime** [2851-68-5] *M* 165.1, *m* 212°, 216°, *pK*<sub>Est(1)</sub>~3.0, *pK*<sub>Est(2)</sub>~10. Crystallise it several times from water or EtOH to give colourless needles. [Lions & Martin *J Am Chem Soc* **79** 2738 1957, *Beilstein* **21** III/IV 4746.]

**Pyridine-2,5-dicarboxylic acid (isocinchomeronic acid)** [100-26-5] *M* 167.1, *m* 254°, (267°, dec), *pK*<sub>1</sub><sup>25</sup> 0.60, *pK*<sub>2</sub><sup>25</sup> 2.49, *pK*<sub>3</sub><sup>25</sup> 5.12. Crystallise it from  $\text{H}_2\text{O}$  or dilute HCl. [Napoli *J Inorg Nucl Chem* **32** 1907 1970, *Beilstein* **22** H 153, **22** I 533, **22** II 105, **22** III/IV 1632, **22/4** V 124.]

**Pyridine-3,4-dicarboxylic acid (cinchomeric acid)** [490-11-9] **M 167.1, m 253-255°, 256°, pK<sub>1</sub><sup>25</sup> 1.50 (0.6), pK<sub>2</sub><sup>25</sup> 2.43 (2.95), pK<sub>3</sub><sup>25</sup> 4.78 (5.07).** Crystallise the acid from H<sub>2</sub>O or dilute aqueous HCl. It has also been purified *via* the *dimethyl ester* which is distilled (**b** 95-100°/1.5mm) and hydrolysed with 3.5N HCl, evaporated and recrystallised. [Armarego & Evans *J Appl Chem* **12** 45 1962, Foye et al. *J Med Chem* **9** 61 1966, *Beilstein* **22** H 155, **22** I 534, **22** II 106, **22** III/IV 1641, **22/4** V 135.]

**Pyridine hydrobromide perbromide (pyridinium bromide perbromide)** [39416-48-3] **M 319.9, m 130° (dec), 132-134°(dec), 135°(dec).** It is a very good brominating agent-liberating one mol of Br<sub>2</sub>. Purify it by recrystallisation from glacial acetic acid (33g from 100ml of AcOH) to give orange-red crystals. [Fieser & Fieser's *Reagents for Organic Chemistry* **1** 967 1967, Englert & McElvain *J Am Chem Soc* **51** 865 1929, *Beilstein* **20** II 103, **20/5** V 181.]

**Pyridine hydrochloride** [628-13-7] **M 115.6, m 144°, b 218°/760mm.** Crystallise the salt from CHCl<sub>3</sub>/EtOAc and wash it with Et<sub>2</sub>O. It is *hygroscopic*. [*Beilstein* **20** H 185, **20** I 57, **20** II 103, **20** III/IV 2230, **20/5** V 180.]

**Pyridine N-oxide** [694-59-7] **M 95.1, m 67°, 68-69°, pK<sub>24</sub> 0.79.** Purify the *N-oxide* by crystallisation from Et<sub>2</sub>O and by vacuum sublimation. The *picrolonate* has **m** 182-184°. [Katritzky *J Chem Soc* 2407 1956, *Beilstein* **20** III/IV 2305, **20/5** V 217.]

**Pyridine 3-sulfonic acid** [636-73-7] **M 159.2, m 365-366°(dec), 357°, pK<sub>25</sub> 2.89 (12% aqueous EtOH), 3.22 (H<sub>2</sub>O)(protonation on N).** Purify the acid by recrystallisation from H<sub>2</sub>O or aqueous EtOH as needles or plates. [pKa: Evans & Brown *J Org Chem* **27** 3127 1962, IR: Arnett & Chawla *J Am Chem Soc* **100** 214 1978.] Its UV in 50% aqueous EtOH has  $\lambda_{\text{max}}$  at 208 and 262nm. The *ammonium salt* has **m** 243° (from H<sub>2</sub>O), the *sulfonyl chloride* has **m** 133-134° (from petroleum ether), the *amide* has **m** 110-111° (from H<sub>2</sub>O), the *hydrochloride* has **m** >300°(dec), and the *N-methyl betaine* has **m** 130° (from H<sub>2</sub>O). [Gastel & Wibaut *Rec Trav Chim Pays Bas* **53** 1031 1934, McElvain & Goese *J Am Chem Soc* **65** 2233 1943, Machek *Monatsh Chem* **72** 77 1938, *Beilstein* **22** I 616, **22** II 309, **22/7** V 552.]

**2-Pyridinethiol (2-mercaptopyridine)** [2637-34-5; 73018-10-7] **M 111.2, m 127.4°, 127-130°, 130-132°, pK<sub>1</sub><sup>20</sup> -1.07, pK<sub>2</sub><sup>20</sup> 9.97.** If impure, dissolve it in CHCl<sub>3</sub>, wash it with dilute AcOH, H<sub>2</sub>O, dry (MgSO<sub>4</sub>), evaporate under reduced pressure and recrystallise the residue from \*C<sub>6</sub>H<sub>6</sub> or H<sub>2</sub>O. *2-Methylmercaptopyridine* (**b** 100-104°/33mm, pK<sub>20</sub> 3.59) was formed by treatment with MeI/NaOH. [Albert & Barlin *J Chem Soc* 2394 1959, Phillips & Shapiro *J Chem Soc* 584 1942, *Beilstein* **21** H 45, **21** III/IV 373, **21/7** V 147.]

**4-Pyridinethiol (4-mercaptopyridine)** [4556-23-4] **M 111.2, m 177°, 179-189°, 186°, pK<sub>1</sub><sup>20</sup> 1.43, pK<sub>2</sub><sup>20</sup> 8.86.** Purify the thiol by dissolving ~45g in boiling H<sub>2</sub>O (100ml) (charcoal), filter and precipitate it by adding 50% aqueous NaOH (~80ml) to pH ~6. Dissolve the precipitate in EtOH, evaporate it to dryness, then crystallise it from boiling EtOH (~100ml, charcoal) to give yellow flat hexagonal plates (**m** 186°). It sublimes readily *in vacuo*. [King & Ware *J Chem Soc* 873 1939.] The *picrate* forms yellow needles from H<sub>2</sub>O with **m** 222°(dec). The *4-methylmercaptopyridine* derivative crystallises from petroleum ether (**m** 47°, also **44-45°** was reported, with a pK<sub>20</sub> of **5.97**) and was prepared by treatment with MeI/NaOH. Its *picrate* has **m** 245° (from H<sub>2</sub>O, MeOH or EtOH). The *N-methyl-4-pyridinethiol* derivative has **m** 168.5-170° (from EtOH), a pK<sub>20</sub> of **1.30** and is soluble in CHCl<sub>3</sub>. [Albert & Barlin *J Chem Soc* 2384 1959, *Beilstein* **21** II 35, **22** III/IV 373, **22/7** V 147.]

**Pyridoxal hydrochloride, pyridoxamine hydrochloride and pyridoxine hydrochloride (vitamin B<sub>6</sub>).** See entries in "Miscellaneous Compounds" in Chapter 7.

**1-(2-Pyridylazo)-2-naphthol (PAN)** [85-85-8] **M 249.3, m 140-142°, 142°, pK<sub>1</sub><sup>30-36</sup> 2.9, pK<sub>2</sub><sup>30-36</sup> 11.2.** Purify PAN by repeated crystallisation from EtOH or MeOH. It can also be purified by sublimation under vacuum. Purity can be checked by TLC using a mixed solvent (petroleum ether/Et<sub>2</sub>O/EtOH; 10:10:1) on a silica gel plate. It has pK<sub>1</sub> 1.9 and pK<sub>2</sub> 12.2 in 20% aqueous ethoxyethanol. It chelates with copper [Pease & Williams *Anal Chem* **31** 1046 1959]. [*Beilstein* **22** I 694, **22** III/IV 7073, **22/4** V 618.]

**4-(2-Pyridylazo)-resorcinol (PAR)** [1141-59-9] M 215.2, m >195°(dec),  $\lambda_{\max}$  415nm,  $\epsilon$  2.59 x 10<sup>4</sup> (pH 6-12), pK<sub>1</sub><sup>25</sup> 2.69 (3.1), pK<sub>2</sub><sup>25</sup> 5.50 (5.8), pK<sub>3</sub><sup>25</sup> 12.5 (11.9). Purify PAR as the sodium salt by recrystallisation from 1:1 EtOH/water. Purity can be checked by TLC using a silica gel plate and a mixed solvent (*n*-BuOH:EtOH:2M NH<sub>3</sub>; 6:2:2). [Beilstein 22 I 694, 22 III/IV 7074, 22/14 V 619.]

**3-(2-Pyridyl)-5,6-diphenyl-1,2,4-triazine** [1046-56-6] M 310.4, m 191-192°. Purify it by repeated recrystallisation from EtOH/dimethylformamide. It is a reagent for estimating Fe(II) and Ru(II). [Chriswell & Schilt *Anal Chem* 46 992 1974, Kamra & Ayers *Anal Chem* 78 423 1975.]

**1-(4-Pyridyl)ethanol** [*R*(+) 27854-88-2; *S*(-) 54656-96-1] M 123.2, m 63-65°, 67-69°, b 138-140°/30mm, 254°/760mm,  $[\alpha]_{\text{D}}^{20}$  *R*(+) +49.8° and *S*(-) -49.8° (c 0.5, EtOH), pK<sub>Est</sub> ~5.4. Purify it by recrystallisation from petroleum ether. The *m* recorded after sublimation was 59.9-60.2°, and 55° after crystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or petroleum ether/\*C<sub>6</sub>H<sub>6</sub>. The (-)-*di-O-benzoyl tartrate salt* has *m* 146-148° (from EtOH). [UV, ORD: Harelli & Samori *J Chem Soc Perkin Trans 2* 1462 1974.] The *racemate* recrystallises from Et<sub>2</sub>O with *m* 74-76°, b 90-94°/1mm. The *picrate* has *m* 125-126° (from \*C<sub>6</sub>H<sub>6</sub>). [Ferles & Attia *Col Czech Chem Commun* 38 611 1973, UV, NMR: Nielson et al. *J Org Chem* 29 2898 1964, Beilstein 21 III/IV 522, 21/2 V 217.]

**5-(3-Pyridyl)-4H-1,2,4-triazole-3-thiol** [1,2-dihydro-5-(3-pyridinyl)-3H-1,2,4-triazole-3-thione] [32362-88-2] M 178.2, , m 285-286°, 296-302°, pK<sub>Est(1)</sub> ~3.0 (py N), pK<sub>Est(2)</sub> ~9 (SH). This thio-1,2,4-triazole was prepared by fusing nicotinoylhydrazide with thiourea at ~150° then at 160-165° (30 minutes), cooled, EtOH was added and the white solid was filtered off [Beyerman et al. *Rec Trav Chim Pays Bas* 73 109 1954]; or by heating 1-nicotinoyl-3-thiosemicarbazide in 5% aqueous NaOH at 100° (4 hours), cooled, acidified with AcOH, and filtered off [Yale & Piala *J Med Chem* 9 42 1966]. It was recrystallised from 95% EtOH, *n*-BuOH/H<sub>2</sub>O, H<sub>2</sub>O, and 0.5% aqueous HCl provides the *hydrochloride*, which is dried *in vacuo* [over P<sub>2</sub>O<sub>5</sub> or H<sub>2</sub>SO<sub>4</sub>]. [Yoshida & Asai *J Pharm Soc Jpn* 74 948 1954, *Chem Abstr* 49 10937 1955, Beilstein 26 IV 2129].

**5-(4-Pyridyl)-4H-1,2,4-triazole-3-thiol** [1,2-Dihydro-5-(4-pyridinyl)-3H-1,2,4-triazole-3-thione] [14910-06-6] M 178.2, , m 301-303° (dec), 305-308° (dec), 308-313°, 320-322° (dec), 321-322° (dec), pK<sub>Est(1)</sub> ~3.5 (py N), pK<sub>Est(2)</sub> ~9 (SH). This thio-1,2,4-triazole was prepared and purified as for the 3-pyridyl-isomer above but by using isonicotinic acid derivatives. It has mild diuretic and natriuretic activity. [Blackman et al. *J Chem Soc C* 661 1967, Beyerman et al. *Rec Trav Chim Pays Bas* 73 109 1954, Yale & Piala *J Med Chem* 9 42 1966, Yoshida & Asai *J Pharm Soc Jpn* 74 948 1954, *Chem Abstr* 49 10937 1955, Beilstein 26 IV 2129]. **5-(4-Pyridyl)-4H-1,2,4-triazole-3-sulfonic acid hydrate** was obtained by oxidising the 3-thiol with saturated aqueous KMnO<sub>4</sub> 25° (1 hour), decomposed with EtOH, filtered, precipitated as the Ag salt, filtered off, and decomposed with 5% HCl, then evaporated, and the residue was recrystallised from H<sub>2</sub>O. It decomposed on heating. [Blackman et al. *J Chem Soc C* 661 1967.]

**$\alpha$ -Pyrone (2H-pyran-2-one, coumalin)** [504-31-4] M 96.1, m 5°, 8-9°, b 103-111°/19-22mm, 110°/26mm, 104°/ 30mm, 115-118°/37mm, 206-207°/atm,  $d_4^{20}$  1.1972,  $n_{\text{D}}^{20}$  1.5298, pK<sup>25</sup> -1.14 (aqueous H<sub>2</sub>SO<sub>4</sub>). Dissolve  $\alpha$ -pyrone in Et<sub>2</sub>O, wash it with brine, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate, distil the residue under vacuum and redistil it. It is a colourless liquid. Its IR has  $\nu_{\max}$  at 1622 and 1752cm<sup>-1</sup> (CHCl<sub>3</sub>). [Zimmermann et al. *Org Synth Coll Vol V* 982 1973, Nakagawa & Saegusa *Org Synth* 56 49 1977, Elderfield *J Org Chem* 6 566 1941.] The *picrate* has *m* 106-107° (from EtOH). [Beilstein 17 H 271, 17 II 305, 17 III/IV 4399, 17/9 V 288.]

**$\gamma$ -Pyrone (4H-pyran-4-one)** [108-97-4] M 96.1, m 32.5-32.6°, 33°, 32-34°, b 88.5°/7mm, 91-91.5°/9mm, 95-97°/13mm, 105°/23mm, 215°/atm, pK<sup>25</sup> 0.10. Purify  $\gamma$ -pyrone by vacuum distillation; the distillate crystallises and is *hygroscopic*. It is non-steam volatile. The *hydrochloride* has *m* 139° (from EtOH), and the *picrate* has *m* 130.2-130.3° (from EtOH or H<sub>2</sub>O). [Mayer *Chem Ber* 90 2362 1957, IR: Jones et al. *Can J Chem* 37 2007 1959, Neelakatan *J Org Chem* 22 1584 1957, Beilstein 17 H 271, 17 I 145, 17 II 305, 17 III/IV 4399, 17/9 V 290.]

**Pyronin G [3,6-bis(dimethylamino)xanthylium chloride]** [92-32-0] **M 302.8, m 250-260°, CI 45005,  $\lambda_{\max}$  522nm,  $pK_{\text{Est}} \sim 7.6$ .** Commercial material may contain a large quantity of zinc. Purify it by dissolving 1g in 50ml of hot water containing 5g NaEDTA. Cool to 0°, filter, evaporate to dryness and the residue is extracted with EtOH. The solution is evaporated to 5-10ml, filtered, and the dye is precipitated by addition of excess of dry diethyl ether. It is centrifuged, and the crystals are washed with dry ether. The procedure is repeated, then the product is dissolved in  $\text{CHCl}_3$ , filtered and evaporated. The dye is stored in a vacuum. [Beilstein 18 H 596, 18 III/IV 7361, 18/10 V 181.]

**Pyrrrol-2,5-dione (maleimide)** [541-59-3] **M 97.1, m 91-93°, 92.6-93°,  $d_{\text{D}}^{105.5}$  1.2493,  $n_{\text{D}}^{110.7}$  1.49256.** Purify it by sublimation in a vacuum. The UV has  $\lambda_{\max}$  at 216 and 280nm in EtOH. [de Wolf & van de Straete *Bull Soc Chim Belg* 44 288 1935, UV: Rondestvedt et al. *J Am Chem Soc* 78 6115 1956, IR: Chiorboli & Mirone *Ann Chim (Rome)* 42 681 1952, Beilstein 21/10 V 3.]

**Pyrrrole** [109-97-7] **M 67.1, m 23.4°, b 66°/80mm, 129-130°/atm,  $d_4^{20}$  0.966,  $n_{\text{D}}^{20}$  1.5097,  $pK_1^{25}$  -4.4 (Protonation on carbon),  $pK_2^{25}$  17.51 (aqueous KOH, H<sub>2</sub> scale).** Dry pyrrole with NaOH,  $\text{CaH}_2$  or  $\text{CaSO}_4$ . Fractionally distil it under reduced pressure from  $\text{CaH}_2$ . Store it under nitrogen as it turns brown in air. Redistil it immediately before use. The *picrate* forms orange-red crystals with **m 69°(dec)**. [Beilstein 20 H 4, 20 I 3, 20 II 3, 20 III/IV 61, 20/5 V 3.]

**1H-Pyrrole-1-propanoic acid (3-[pyrrol-1-yl]propionic acid)** [89059-06-3] **M 139.2, m 59-64°, 62.5°,  $pK_{\text{Est}} \sim 4.5$ .** Recrystallise the acid from petroleum ether (b 80-100°) and dry it *in vacuo*. The *ethyl ester* has **b 122°/23mm**. The *amide* forms colourless needles from  $^*\text{C}_6\text{H}_6$  with **m 81°** and is soluble in cold  $\text{H}_2\text{O}$ . [Clemo & Ramage *J Chem Soc* 49 1931, Jefford & Johncock *Helv Chim Acta* 66 2666 1983.]

**Pyrrrolidine** [123-75-1] **M 71.1, b 87.5-88.5°,  $d_4^{20}$  0.860,  $n_{\text{D}}^{20}$  1.443,  $pK^{25}$  11.31.** Dry pyrrolidine with BaO or sodium, then fractionally distil it, under  $\text{N}_2$ , through a Todd column packed with glass helices. [Beilstein 20 H 159, 20 I 36, 20 II 79, 20 III/IV 2072, 20/1 V 162.]

**Pyrrrolidine-1-carbodithioic acid ammonium salt** [5108-96-3] **M 164.3, m 128-130°,  $pK^{25}$  3.25 (free acid).** Purify the salt by recrystallising twice by dissolving in MeOH and adding  $\text{Et}_2\text{O}$ . It has also been purified by recrystallisation from EtOH. It is used for the precipitation of As, Bi, Cd, Co, Cu, Fe, Mn, Ni, Pb, Sb, Su, V and Zn from acidic solutions [Kinrade & van Loon *Anal Chem* 46 1894 1974]. The *ethyl ester* has **b 87-89°/0.01mm**. [Synthesis and Polarography: Kitagawa & Taku *Bull Chem Soc Jpn* 64 2151 1973, Malissa & Schöffmann *Mikrochim Acta* 187 1955, Beilstein 20 III/IV 195, 20/1 V 331.]

**Quercetin (2H<sub>2</sub>O) (3,3',4',5,6-pentahydroxyflavone)** [6151-25-3 (2H<sub>2</sub>O); 117-39-3 (anhydrous)] **M 338.3, m ca 315°(dec), 317.4-317.9°(dec), (phenolic  $pK_s$  7—10).** Crystallise the flavone from aqueous EtOH and dry at 100°. It also sublimes in a vacuum. Its IR has  $\nu_{\max}$  at 2.5 $\mu$  and 16.7 $\mu$  (KBr). It complexes with  $\text{Cu}^{2+}$ ,  $\text{Al}^{3+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Ge}^{4+}$ ,  $\text{Zn}^{2+}$ ,  $\text{Ti}^{4+}$ ,  $\text{Zr}^{4+}$ ,  $\text{Th}^{4+}$ ,  $\text{UO}^{4+}$ . [Beilstein 18 H 242, 18 I 242, 18 II 236, 18 III/IV 3470, 18/5 V 494.]

(±)-**Quinacrine [Atebrine, Mepacrine, 3-chloro-9(4-diethylamino-1-methyl)butylamino-7-methoxy-acridine] dihydrochloride.** [69-05-6] **M 472.9, m 248-250°(dec),  $pK_1^{30}$  -6.49 (aq  $\text{H}_2\text{SO}_4$ ),  $pK_2^{30}$  7.73 (ring  $\text{NH}^+$ ),  $pK_3^{30}$  10.18 ( $\text{Et}_2\text{N}$ ).** It crystallises from  $\text{H}_2\text{O}$  (solubility is 2.8% at room temperature) as yellow crystals. It is slightly soluble in MeOH and EtOH. The *free base* crystallises from  $\text{Me}_2\text{CO}$  or petroleum ether with **m 86-88°**, or aqueous EtOH with 85-87.5°. The *bismethiodide* has **m 224°** (from MeOH/EtOAc/ $\text{Et}_3\text{N}$ ), and the *picrate* has **m 207-208°(dec)** when crystallised from  $\text{Me}_2\text{CO}$ /EtOH. It is an antimalarial, antiprotozoal and intercalates DNA. [Wolfe *Antibiot* 3 (Springer-Verlag) 203 1975, Beilstein 22 III/IV 6247, 22/12 V 235.]

**Quinaldic (quinoline-2-carboxylic) acid** [93-10-7] **M 173.2, m 156-157°,  $pK_1^{25}$  1.45,  $pK_2^{25}$  2.49 (2.97).** Crystallise quinaldic acid from  $^*\text{C}_6\text{H}_6$  or AcOH. It is used for the estimation of many metals. The *methyl ester* has **m 86-87°** (from hexane) and  $pK^{25}$  1.76. [Chaudhuri et al. *Frez Z Anal Chem* 281 361 1976, Beilstein 22 H 71, 22 II 55, 22 III/IV 1149, 22/3 V 183.]

**Quinazoline** [253-82-7] **M** 130.2, **m** 48.0-48.5°, **b** 120-121°/17-18mm, **pK**<sub>1</sub><sup>20</sup> -4.51 (aqueous H<sub>2</sub>SO<sub>4</sub>, anhydrous dication), **pK**<sub>2</sub><sup>20</sup> 2.01 (anhydrous monocation), **pK**<sub>3</sub><sup>20</sup> 4.3 (equilibrium with 3,4-hydrated species), **pK**<sub>4</sub><sup>20</sup> 12.1 (hydrated anion). Purify quinazoline by passage through an activated alumina column in \*C<sub>6</sub>H<sub>6</sub> or petroleum ether (b 40-60°). Distil it under reduced pressure, sublime it under vacuum and crystallise it from petroleum ether. The *picrate* has **m** 188-189° (from MeOH). [Armarego *J Appl Chem* **11** 70 1961, Armarego *Quinazolines, Fused Pyrimidines Part I* Brown Ed, Wiley-Interscience 1967, Brown *Quinazolines Supplement I* Taylor Ed, Wiley-Interscience 1996, ISBN 0-471-14565-3; for covalent hydration see Albert & Armarego *Adv Heterocycl Chem* **4** 1 1965, *Beilstein* **23** H 175, **23** II 177, **23** III/IV 1221.]

**S(+)-Quinidine** [56-54-2] **M** 324.4, **m** 171°,  $[\alpha]_{546}^{20}$  +301.1° (CHCl<sub>3</sub> containing 2.5% (v/v) EtOH), **pK**<sub>1</sub><sup>15</sup> 4.13, **pK**<sub>2</sub><sup>15</sup> 8.77. Crystallise it from \*C<sub>6</sub>H<sub>6</sub> or dry CHCl<sub>3</sub>/petroleum ether (b 40-60°), discarding the initial, oily crop of crystals. Dry it under vacuum at 100° over P<sub>2</sub>O<sub>5</sub>. It has been used as a chiral catalyst [Wynberg & Staring *J Am Chem Soc* **104** 166 1982, *J Org Chem* **50** 1977 1985]. [*Beilstein* **23** H 506, **23** I 164, **23** II 414, **23** III/IV 3261, **23/13** V 395.]

**8S,9R-Quinine** [130-95-0] **M** 324.4, **m** 177°(dec),  $[\alpha]_{546}^{20}$  -160° (c 1, CHCl<sub>3</sub>), **pK**<sub>1</sub><sup>20</sup> 4.13 (quinoline N), **pK**<sub>2</sub><sup>20</sup> 8.52 (piperidine N). Crystallise the quinine from absolute EtOH. It has been used as a chiral catalyst (see previous entry). [*Beilstein* **23** H 511, **23** I 166, **23** II 416, **23** III/IV 3265, **23/13** V 395.]

**Quinine bisulfate** [6183-68-2 (7H<sub>2</sub>O), 549-56-4 (anhydrous)] **M** 422.4, **m** 160° (anhydrous). Crystallise the bisulfate from 0.1M H<sub>2</sub>SO<sub>4</sub>, and recrystallise it from water to give the *heptahydrate*. [*Beilstein* **23** III/IV 3270, **23/13** V 396.]

**Quinine sulfate (2H<sub>2</sub>O)** [6119-70-6 (H<sub>2</sub>O), 804-63-7 (anhydrous)] **M** 783.0, **m** 205° (219°). Crystallise it from water and dry it at 110°. [*Beilstein* **23** H 522, **23** I 168, **23** II 420, **23** III/IV 3269, **23** V 396.]

**Quinoline** [91-22-5] **M** 129.2, **m** -20°, -16°, **b** 113-114°/17mm, 236°/758mm, **d**<sub>4</sub><sup>20</sup> 1.0937, **n**<sub>D</sub><sup>20</sup> 1.625, **pK**<sup>25</sup> 4.80 (4.93). Dry quinoline with Na<sub>2</sub>SO<sub>4</sub> and distil it from zinc dust in a vacuum. It has also been dried by boiling with acetic anhydride, then fractionally distilled. Calvin and Wilmarth [*J Am Chem Soc* **78** 1301 1956] cooled redistilled quinoline in ice and added enough HCl to form its hydrochloride. Diazotization removed aniline, the diazo compound being broken down by warming the solution to 60°. Non-basic impurities were removed by ether extraction. Quinoline was then liberated by neutralising the hydrochloride with NaOH, then dried with KOH and fractionally distilled at low pressure. Addition of cuprous acetate (7g/L of quinoline) and shaking under hydrogen for 12 hours at 100° removed impurities due to the nitrous acid treatment. Finally the hydrogen was pumped off, and the quinoline was distilled. Other purification procedures depend on conversion to the *phosphate* (**m** 159°, precipitated from MeOH solution, filtered, washed with MeOH, then dried at 55°) or the *picrate* (**m** 201°) which, after recrystallisation, were reconverted to quinoline.

The method using the *picrate* [Packer et al. *J Am Chem Soc* **80** 905 1958] is as follows: quinoline is added to picric acid dissolved in the minimum volume of 95% EtOH, giving yellow crystals which are washed with EtOH, air-dried and crystallised from acetonitrile. These are dissolved in dimethyl sulfoxide (previously dried over 4A molecular sieves) and passed through a basic alumina column, onto which the picric acid is adsorbed. The free base in the effluent is extracted with *n*-pentane and distilled under vacuum. Traces of solvent can be removed by vapour-phase chromatography. [Moonaw & Anton *J Phys Chem* **80** 2243 1976.] The ZnCl<sub>2</sub> and dichromate complexes have also been used [Cumper et al. *J Chem Soc* 1176 1962]. [*Beilstein* **20** H 339, **20** I 134, **20** II 222, **20** III/IV 3334, **20/7** V 276.]

**2-Quinolinealdehyde** [5470-96-2] **M** 157.2, **m** 71°, **pK**<sub>Est</sub> ~3.3. Distil it in steam and recrystallise it from H<sub>2</sub>O. Protect it from light. The *semicarbazone* has **m** 254° (from aqueous EtOH), and the *picrate* has **m** 197-199°. [*Beilstein* **21** H 322, **21** III/IV 4034, **21/8** V 442.]

**8-Quinolinecarboxylic acid** [86-59-9] **M** 173.2, **m** 186-187.5°, **pK**<sub>1</sub><sup>25</sup> 1.82, **pK**<sub>2</sub><sup>25</sup> 6.87. Crystallise the acid from water, aqueous EtOH, EtOH or \*C<sub>6</sub>H<sub>6</sub>. The *ethyl ester* has **m** 45° and **b** 194-197°/13mm. [*Beilstein* **22** H 81, **22** III/IV 1200, **22/3** V 217.]

Quinoline ethiodide (1-ethylquinolinium iodide) [634-57-5] M 285.1, m 158-159°. Crystallise it from aqueous EtOH or EtOH peroxide free Et<sub>2</sub>O. [Beilstein 20 H 352, 20 I 139, 20 II 231, 20 III/IV 3357, 20/7 V 276.]

Quinoxaline [91-19-0] M 130.2, m 28° (anhydrous), 37°(H<sub>2</sub>O), b 108-110°/0.1mm, 140°/40mm, pK<sub>1</sub><sup>20</sup> -5.52 (-5.8, dication), pK<sub>2</sub><sup>20</sup> 0.56 (c 0.5), 0.72 (c 1.0) (monocation). Crystallise quinoxaline from petroleum ether. It crystallises as the *monohydrate* on addition of water to a petroleum ether solution. It has UV with  $\lambda_{\max}$  at 242 and 331nm (H<sup>+</sup> -2); 234 and 316nm (pH 7.1). The *picrate* has m 161-162°. [Albert & Phillips *J Chem Soc* 1294 1956, *Beilstein* 23 H 176, 23 II 177, 23 III/IV 1226, 23/7 V 135.]

Quinoxaline-2,3-dithiol [1199-07-7] M 194.1, m 345°(dec), pK<sub>1</sub> 6.9, pK<sub>2</sub> 9.9. Purify the dithiol by repeated dissolution in alkali and re-precipitation by acetic acid. It complexes with Ag<sup>+</sup>, Cd<sup>2+</sup>, Pb<sup>2+</sup>, Bi<sup>3+</sup> and Ni<sup>2+</sup> in aqueous NH<sub>3</sub>. [*Beilstein* 24 III/IV 1428.]

Quinuclidine (1-azabicyclo[2.2.2]octane) [100-76-5] M 111.2, m 158°(sublimes), pK<sup>25</sup> 10.95. Crystallise it from diethyl ether. The *hydrochloride* has m 364-365°(dec) (from EtOH or n-BuOH), and the *picrate* has m 225° (from aqueous EtOH). [*Beilstein* 20 H 144, 20 II 71, 20 III/IV 1966, 20/4 V 335.]

**Rescinnamine** (Anaprel, Apoterin, a  $\beta$ -carboline alkaloid) [24815-24-5] M 634.7, m 238-239°(vacuum), 240°,  $[\alpha]_{\text{D}}^{20}$  -97° (c 1, CHCl<sub>3</sub>), pK<sub>EST(1)</sub> <0 (carbazole N), pK<sub>EST(2)</sub> ~7.0 (quinolizidine N), pK<sup>25</sup> 6.4 (75% aqueous HCONMe<sub>2</sub>). Crystallise it from \*benzene, MeOH or aqueous Me<sub>2</sub>CO. The *hydrochloride* has m 232°(dec) (from MeOH) and  $[\alpha]_{\text{D}}^{25}$  -74° (MeOH). It is an antihypertensive. [Klohs et al. *J Am Chem Soc* 77 2241 1955, *Beilstein* 25 III/IV 1323.]

(-)-Reserpine acid [83-60-3] M 400.5, m 241-243°,  $[\alpha]_{\text{D}}^{24}$  -70° (c 1, H<sub>2</sub>O), pK<sub>EST(1)</sub> <0 (carbazole N), pK<sub>EST(2)</sub> -4.0 (C<sub>2</sub>H<sub>4</sub>), pK<sub>EST(3)</sub> -7.4 (quinolizidine N), pK<sub>1</sub> 6.2, pK<sub>2</sub> 8.2 (66% aqueous HCONMe<sub>2</sub>). Crystallise the acid from MeOH. It has UV with  $\lambda_{\max}$  at 222, 268 and 294nm (EtOH). The *hydrochloride 0.5H<sub>2</sub>O* has m 257-259° (from EtOH/Et<sub>2</sub>O),  $[\alpha]_{\text{D}}^{23}$  -81° (H<sub>2</sub>O). [Dorfman et al. *Helv Chim Acta* 77 59 1954, *Beilstein* 25 III/IV 1305.]

Reserpine [50-55-5] M 608.7, m 262-263°,  $[\alpha]_{\text{D}}^{20}$  -148° (c 1, CHCl<sub>3</sub>), pK<sub>EST(1)</sub> <0 (carbazole N), pK<sub>2</sub> 6.6 (7.4)(quinolizidine N). Crystallise reserpine from aqueous Me<sub>2</sub>CO or Et<sub>2</sub>O. [Woodward et al. *Tetrahedron* 2 155 1958, *Beilstein* 25 III/IV 1319.]

Rhamnetin (3,3'-4',5-tetrahydroxy-7-methoxy flavone, 7-methyl quercetin) [90-19-7] M 316.3, m >300°(dec), several phenolic pKs -7-10.5. Crystallise rhamnetin from EtOH (m 292-293°), aqueous EtOH (m 294-296°) or MeOH (m 290-294°), or Me<sub>2</sub>CO/MeOH. [Kuhn & Low *Chem Ber* 77 211 1944, Jurd *J Am Chem Soc* 80 5531 1958.] The *tetra-acetate* has m 189-190° (from Me<sub>2</sub>CO/MeOH). [*Beilstein* 18 H 245, 18 II 237, 18 III/IV 3474.]

Rhodamine 3B chloride [3,5-bis-(diethylamino)-9-(2-carboxyphenyl)xanthylium chloride] [81-88-9] M 479.0, m 210-211°(dec), CI 45170,  $\lambda_{\max}$  543nm, {Free base [509-34-2] CI 749}, pK<sup>25</sup> 5.53. Major impurities are partially dealkylated compounds not removed by recrystallisation. Purify the dye by chromatography, using ethyl acetate/isopropanol/ammonia (conc) 9:7:4, R<sub>f</sub> 0.75 on Kieselgel G). It has also been crystallised from a concentrated solution in MeOH by slow addition of dry diethyl ether; or from EtOH containing a drop of conc HCl by slow addition of ten volumes of dry diethyl ether. The solid is washed with ether and air dried. The dried material has also been extracted with \*benzene to remove oil-soluble material prior to recrystallisation. Store it in the dark. [*Beilstein* 18 II 486, 18 III/IV 8246, 19/8 V 669.]

Rhodamine 6G [Basic Red 1, 3,5-bis-(ethylamino)-9-(2-ethoxycarbonylphenyl)-2,7-dimethylxanthylium chloride] [989-38-8] M 479.3, CI 45160,  $\lambda_{\max}$  524nm, pK<sup>25</sup> 5.58. Crystallise the dye from MeOH or EtOH, and dry it in a vacuum oven. [*Beilstein* 18 III/IV 8244, 18/12 V 283.]

**Rhodanine (2-mercaptothiazolidin-4-one)** [141-84-4] **M 133.2, m 168.5° (capillary), pK<sup>20</sup> 5.18.** Crystallise rhodanine from glacial acetic acid or water. It is used to estimate Ag and gallic acid [Thies & Fischer *Microchimica Acta* 809 1973]. [Beilstein 27 H 242, 27 I 309, 27 II 288, 27 III/IV 3188.]

**Riboflavin, riboflavin-5'-phosphate (Na salt, 2H<sub>2</sub>O).** See entries in “Miscellaneous Compounds” and for **ribonucleic acids** see entry in “Proteins, Enzymes, DNA and RNA” in Chapter 7.

**Saccharin (1,2-benzisothiazol-3(2H)-one 1,1-dioxide, *o*-benzoic acid sulfimide)** [81-07-2] **M 183.2, m 227-229°, 229°, 228.8-229.7°, pK<sub>1</sub><sup>25</sup> 1.31, pK<sub>2</sub><sup>25</sup> 12.8.** Purify saccharin by recrystallisation from Me<sub>2</sub>CO [solubility 7.14% at 0°, 14.4% at 50°], or aqueous isoPrOH to give a fluorescent solution. It sublimes *in vacuo*. It is an artificial sweetener and is 500 times sweeter than sucrose. [DeGarmo et al. *J Am Pharm Assoc (Sci Ed)* 41 17 1952, Beilstein 27 H 168, 870, 27 I 266, 27 II 214, 27 III/IV 2649.]

**Safranine O (Safranine T, 3,7-diamino-2,8-dimethyl-5-phenylphenazinium chloride)** [477-73-6] **M 350.9, λ<sub>max</sub> 530nm, pK<sup>25</sup> 6.4.** Crystallise it from \*benzene/MeOH (1:1) or water. Dry it *in vacuo* over H<sub>2</sub>SO<sub>4</sub>. It has UV: λ<sub>max</sub> at 520nm (H<sub>2</sub>O) and 530nm (EtOH). [Beilstein 25 H 403, 25 I 657, 25 III/IV 3056.]

**Safrole (5-allyl-1,3-benzodioxole, 4-allyl-1,2-methylenedioxybenzene)** [94-59-7] **M 162.1, m~ 11°, b 69-70°/1.5mm, 104-105°/6mm, 231.5-232°/atm, 235-237°/atm, d<sub>4</sub><sup>20</sup> 1.0993, n<sub>D</sub><sup>20</sup> 1.53738.** Safrole has been purified by fractional distillation, although it has also been recrystallised from low boiling petroleum ether at low temperatures. [IR: Briggs et al. *Anal Chem* 29 904 1957, UV: Patterson & Hibbert *J Am Chem Soc* 65 1962 1943.] The *maleic anhydride adduct* forms yellow crystals from toluene **m 257°** [Hickey *J Org Chem* 13 443 1948], and the *picrate* forms orange-red crystals from CHCl<sub>3</sub> [Baril & Magrdichian *J Am Chem Soc* 58 1415 1936]. [Beilstein 19 H 39, 19 I 617, 19 II 29, 19 III/IV 275, 19/1 V 553.]

**Scopolamine (hyoscine, atropine, 6β,7β-epoxy-3-α-tropanyl *S*-(-)-tropate)** [51-34-3] **M 321.4, m 59°, [α]<sub>D</sub><sup>20</sup> -18° (c 5, EtOH), -28° (c 2, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> -30° (c 5, CHCl<sub>3</sub>), pK<sup>25</sup> 7.55.** Crystallise it from \*benzene/petroleum ether, EtOH or H<sub>2</sub>O. It is polymorphic with **m 165-166° and 190-191°.** The *racemate* has **m 56-57° (H<sub>2</sub>O), 37-38° (2H<sub>2</sub>O),** syrup (anhydrous), *l* and *d* isomers can separate as syrups when anhydrous. [Beilstein 27 H 99, 102, 27 I 247-248, 27 II 43-44, 27 III/IV 1790.]

**Scopoletin (7-hydroxy-6-methoxycoumarin)** [92-61-5] **M 192.2, m 206°, 208-209°, pK<sup>25</sup> 8.96 (70%aqueous EtOH).** Crystallise it from water, acetic acid or \*C<sub>6</sub>H<sub>6</sub>/MeOH. It is dimorphic with a second **m** at 193-195°. It sublimes at 120-130°/12mm. [Beilstein 18 H 99, 18 I 348, 18 II 68, 18 III/IV 1323, 18/3 V 203.]

**Secobarbital (5-allyl-5-1'-methylbutylbarbituric acid)** [76-73-3] **M 260.3, m 100°, pK<sub>1</sub><sup>25</sup> 8.08, pK<sub>2</sub><sup>25</sup> 12.6.** It is purified by dissolving the *sodium salt* [309-43-3] in 10% HCl which precipitates the acid form that is extracted into Et<sub>2</sub>O. The extract is dried (Na<sub>2</sub>SO<sub>4</sub>) and evaporated. The residue is then purified by repeated crystallisation from CHCl<sub>3</sub>. [Buchet & Sandorfy *J Phys Chem* 88 3274 1984, Beilstein 24 III/IV 2013.]

**Serotonin creatinine sulfate (H<sub>2</sub>O)** [971-74-4] **M 405.4, m 220°(dec), pK<sub>1</sub> 10.1, pK<sub>2</sub> 11.1, pK<sub>3</sub> 18.25 (NH) for serotonin, pK<sup>25</sup> 4.9 for creatinine.** It crystallises as the monohydrate from water. [Beilstein 22 III/IV 5667, 22/12 V 18.]

**Sinomenine hydrochloride (7,8-didehydro-4-hydroxy-3,7-dimethoxy-17-methyl-9α,13α,14α-morphan-6-one HCl)** [6080-33-7] **M 365.9, m 231°, [α]<sub>D</sub><sup>17</sup> -83° (c 4, H<sub>2</sub>O), pK<sub>Est(1)</sub>~10.0 (N), pK<sub>Est(2)</sub>~10.4 (OH).** Crystallise the salt from water (1g/1.5ml) or EtOH. The *free base* [115-53-7] **M 329.4,** has **m 161°** (from EtOH) (and again at 182°) after crystallisation from \*C<sub>6</sub>H<sub>6</sub>, and [α]<sub>D</sub><sup>20</sup> -78.9° (c 1, EtOH). The *picrate* has **m 159-162°(dec)** (from H<sub>2</sub>O). [Beilstein 21 II 470, 21 III/IV 6670.]

**(-)-Sparteine sulfate pentahydrate** [6160-12-9] **M 422.5, m loses H<sub>2</sub>O at 100° and turns brown at 136° (dec), [α]<sub>D</sub><sup>20</sup> -22° (c 5, H<sub>2</sub>O), [α]<sub>D</sub><sup>21</sup> -16° (c 10, EtOH for free base), pK<sub>1</sub><sup>20</sup> 2.24, pK<sub>2</sub><sup>20</sup> 9.46.** Recrystallise the

sulfate from aqueous EtOH or H<sub>2</sub>O, although the solubility in the latter is high. The *free (-)-base* has **b** 173°/8mm and is steam volatile but resinifies in air. The *dipicrate* forms yellow needles from EtOH/Me<sub>2</sub>CO, **m** 205-206° [Clemo et al. *J Chem Soc* 429 1931; see also Bolnmann & Schuman *The Alkaloids* (Ed Manske) Vol 9 175 1967]. The *free (±)-base* has **m** 71-72.5° [van Tamelen & Foltz *J Am Chem Soc* 82 2400 1960]. [Beilstein 23 II 99, 23 III/IV 18.]

**(-)-Strychnine** [57-24-9] **M** 334.4, **m** 268°, 278-279.5°,  $[\alpha]_{546}^{20}$  -139° (c 1, CHCl<sub>3</sub>), **pK<sub>1</sub><sup>20</sup>** 2.50, **pK<sub>2</sub><sup>20</sup>** 8.2, **pK<sub>1</sub><sup>20</sup>** 7.37 [80% aqueous MeO(CH<sub>2</sub>)<sub>2</sub>OH]. It crystallises from CHCl<sub>3</sub>/Et<sub>2</sub>O and sublimes at 125°/0.01mm. It can also be purified by conversion to the *hydrochloride* [**m** 275-295° (dec),  $[\alpha]_{D}^{20}$  -44° (0.03N HCl)] with aqueous HCl, then neutralisation with ammonia. [Beilstein 27 II 723, 27 III/IV 7530.] It is POISONOUS.

**Sulfamethazine (Sulfadimidine, N'-[4,6-dimethyl-2-pyrimidinyl]sulfanilamide)** [57-68-1] **M** 278.3, **m** 178-179°, 198-200°, 205-207°, **pK<sub>1</sub>** 2.65, **pK<sub>2</sub>** 7.4. Crystallise it from dioxane or aqueous dioxane. [Caldwell et al. *J Am Chem Soc* 63 2188 1941, Roblin et al. *J Am Chem Soc* 64 567 1942, Beilstein 25 III/IV 2215.]

**Sulfapyridine** [144-83-2] **M** 349.2, **m** 193°, **pK<sup>20</sup>** 8.64. Crystallise sulfapyridine from 90% acetone and dry it at 90°. Its solubility in Me<sub>2</sub>CO, EtOH and H<sub>2</sub>O is 1.5%, 0.22% and 0.02%, respectively. [Winterbottom *J Am Chem Soc* 62 160 1940, Beilstein 22 III/IV 3978.]

**2-Sulfobenzoic cyclic anhydride (2,1-benzoxathiazol-3-one-1,1-dioxide)** [81-08-3] **M** 184.2, **m** 126-127°, 129.5°, 130°, **b** 184-186°/18mm. The anhydride is purified by distillation in a vacuum and readily solidifies to a crystalline mass on cooling. [Heitman *J Am Chem Soc* 34 1594 1912.] Alternatively, purify it by dissolving it in the minimum volume of toluene and refluxing for 2 hours using a Dean-Stark trap. Evaporate under reduced pressure and distil the anhydride at 18mm. It is then recrystallised three times from its own weight of dry \*C<sub>6</sub>H<sub>6</sub>. It is sensitive to moisture and should be stored in the dark in a dry atmosphere. The *O-methyloxime* has **m** 110-112° [Levy *Tetrahedron Lett* 3289 1972]. If the sample has hydrolysed extensively (presence of OH band in the IR) then treat with an equal bulk of SOCl<sub>2</sub>, reflux it for 3 hours (CaCl<sub>2</sub> tube), evaporate and distil the residue in a vacuum, then recrystallise it from \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub> (EtOH free by passing through Al<sub>2</sub>O<sub>3</sub>, or standing over CaCl<sub>2</sub>). [Clarke & Dreger *Org Synth Coll Vol I* 495 1941.] It is used for modifying ζ-amino functions of lysyl residues in proteins [Bagree et al. *FEBS Lett* 120 275 1980]. [Beilstein 19 I 659, 19 II 137, 19 III/IV 1641, 19/4 V 215.]

**Sulfolane (tetramethylenesulfone)** [126-33-0] **M** 120.2, **m** 28.5°, **b** 153-154°/18mm, 285°/760mm, **d<sub>4</sub><sup>20</sup>** 1.263, **n<sub>D</sub><sup>30</sup>** 1.4820. It is prepared commercially by a Diels-Alder reaction of between 1,3-butadiene and sulfur dioxide, followed by Raney nickel hydrogenation. The principal impurities are water, 3-sulfolene, 2-sulfolene and 2-isopropyl sulfolanyl ether. It is dried by passage through a column of molecular sieves. Distil it under reduced pressure through a column packed with stainless steel helices. Again dry it with molecular sieves and distil. [Cram et al. *J Am Chem Soc* 83 3678 1961, Coetzee *Pure Appl Chem* 49 211 1977.] Alternatively, it is stirred at 50°, and small portions of solid KMnO<sub>4</sub> are added until the colour persists during 1 hour. Dropwise addition of MeOH then destroys the excess KMnO<sub>4</sub>; the solution is filtered, freed from potassium ions by passage through an ion-exchange column and dried under vacuum. It has also been distilled in a vacuum from KOH pellets. It is *hygroscopic*. [See Sacco et al. *J Phys Chem* 80 749 1976, *J Chem Soc, Faraday Trans 1* 73 1936 1977, 74 2070 1978, *Trans Faraday Soc* 62 2738 1966.] Coetzee has reviewed the methods of purification of sulfolane, and also the removal of impurities. [Coetzee in *Recommended Methods of Purification of Solvents and Tests for Impurities*, Coetzee Ed. Pergamon Press, 1982, Beilstein 17 I 5, 17 III/IV 37, 17/1 V 39.]

**2,2':6',2''-Terpyridyl** [1148-79-4] **M** 233.3, **m** 91-92°, **pK<sub>1</sub><sup>23</sup>** 2.64, **pK<sub>2</sub><sup>23</sup>** 4.33. Crystallise it from diethyl ether, toluene or from petroleum ether, then aqueous MeOH, followed by sublimation in a vacuum at 90°. It is used for estimating Ag and Ru. [Kamra et al. *Anal Chim Acta* 81 177 1976, Beilstein 26 III/IV 258.]

**Terthiophene (2,5-di[thienyl]thiophene; α-terthienyl)** [1081-34-1] **M** 248.4, **m** 94-95.5°, 94-96°. Possible impurities are bithienyl and polythienyls. Suspend it in H<sub>2</sub>O and steam distil it to remove bithienyl.

The residue is cooled and extracted with  $\text{CHCl}_3$ , dried ( $\text{MgSO}_4$ ), filtered, evaporated and the residue chromatographed on  $\text{Al}_2\text{O}_3$  using petroleum ether/3%  $\text{Me}_2\text{CO}$  as eluent. The terphenyl zone is then eluted from the  $\text{Al}_2\text{O}_3$  with  $\text{Et}_2\text{O}$ , the extract is evaporated and the residue is recrystallised from  $\text{MeOH}$  (40ml per g). The platelets are washed with cold  $\text{MeOH}$  and dried in air. [UV: Sease & Zechmeister *J Am Chem Soc* **69** 270 1947; Uhlenbroek & Bijloo *Rec Trav Chim Pays Bas* **79** 1181 1960.]

It has also been recrystallised from  $\text{MeOH}$ ,  $^*\text{C}_6\text{H}_6$ , petroleum ether or  $\text{AcOH}$ . [UV: Zechmeister & Sease *J Am Chem Soc* **69** 273 1947, Steinkopf et al. *Justus Liebigs Ann Chem* **546** 180 1941.] It is a phototoxic nematocide [Cooper & Nitsche *Bioorg Chem* **13** 36 1985, Chan et al. *Phytochem* **14** 2295 1975]. [*Beilstein* **19** III/IV 4763, **19/9** V 226.]

**2,4,5,6-Tetraaminopyrimidine sulfate** [5392-28-9] **M 238.2, m 255° (dec), >300°, >350° (dec), pK<sup>20</sup> 6.82.** Purify the salt by recrystallisation from  $\text{H}_2\text{O}$ , 2N  $\text{H}_2\text{SO}_4$  (20 parts, 67% recovery) or 0.1N  $\text{H}_2\text{SO}_4$  (40 parts, 62% recovery), and dried in air. [UV: Konrad & Pfeleiderer *Chem Ber* **103** 722 1970, Malletta et al. *J Am Chem Soc* **69** 1814 1947, Cavalieri et al. *J Am Chem Soc* **70** 3875 1948, *Beilstein* **25** H 423, **25** III/IV 3106.]

**1,4,8,11-Tetraazacyclotetradecane (cyclam)** [295-37-4] **M 200.33, m 173° (closed capillary and sublimes at 125°), 183-185°, 185°, pK<sub>Est(1)</sub> ~3.8, pK<sub>Est(2)</sub> ~6.0, pK<sub>Est(3)</sub> ~9.0, pK<sub>Est(4)</sub> ~9.6.** Purify cyclam by recrystallisation from dioxane (white needles), and it sublimes above 120°. It has been distilled, **b 132-140°/4-8mm**. It forms complexes with metals and gives a sparingly soluble *nitrate salt*, **m 205°(dec)**, which crystallises from  $\text{H}_2\text{O}$  and is dried at 150°. [UV: Bosnich et al. *Inorg Chem* **4** 1102 1963, van Alphen *Rec Trav Chim Pays Bas* **56** 343 1937, *Beilstein* **26** III/IV 1647.]

**Tetrabenazine (2-oxo-3-isobutyl-9,10-dimethoxy-1,2,3,4,6,7-hexahydro-11bH-benzo[a]-quinolizine)** [58-46-8] **M 317.4, m 127-128°, pK<sub>Est</sub> ~8.** Crystallise it from  $\text{MeOH}$ . The *hydrochloride* has **m 208-210°**, and the *oxime* has **m 158° (from EtOH)**. [*Beilstein* **21** III/IV 6488.]

**2,3,4,6-Tetrachloropyridine** [14121-36-9] **M 216.9, m 74-75°, (? 37-38°), b 130-135°/16-20mm, 248.5-249.5°/760mm, pK<sub>Est</sub> ~ -5.7.** Crystallise it from 50%  $\text{EtOH}$  and/or distil it. The *N-oxide* has **m 210° (from EtOH/CHCl<sub>3</sub>)**. [Chivers & Suschitzky *J Chem Soc* 2870 1971, *Beilstein* **20** III/IV 6488, **20/5** V 421.]

**Tetrahydrofuran (oxalane, THF)** [109-99-9] **M 72.1, b 25°/176mm, 66°/760mm, d<sub>4</sub><sup>20</sup> 0.889, n<sub>D</sub><sup>20</sup> 1.4070, pK<sup>25</sup> -2.48 (aqueous  $\text{H}_2\text{SO}_4$ ).** It is obtained commercially by catalytic hydrogenation of furan from pentosan-containing agricultural residues. It was purified by refluxing with, and distilling from  $\text{LiAlH}_4$  which removes water, peroxides, inhibitors and other impurities [Jaeger et al. *J Am Chem Soc* **101** 717 1979]. Peroxides can also be removed by passage through a column of activated alumina, or by treatment with aqueous ferrous sulfate and sodium bisulfate, followed by solid  $\text{KOH}$ . In both cases, the solvent is then dried and fractionally distilled from sodium. Lithium wire or vigorously stirred molten potassium have also been used for this purpose.  $\text{CaH}_2$  has also been used as a drying agent.

Several methods are available for obtaining the solvent almost anhydrous. Ware [*J Am Chem Soc* **83** 1296 1961] dried it vigorously with sodium-potassium alloy until a characteristic blue colour was evident in the solvent at Dry-ice/cellosolve temperatures. The solvent is kept in contact with the alloy until distilled for use. Worsfold and Bywater [*J Chem Soc* 5234 1960], after refluxing and distilling from  $\text{P}_2\text{O}_5$  and  $\text{KOH}$ , in turn, refluxed the solvent with sodium-potassium alloy and fluorenone until the green colour of the disodium salt of fluorenone was well established. [Alternatively, instead of fluorenone, benzophenone, which forms a blue ketyl, can be used.] The tetrahydrofuran was then fractionally distilled, degassed and stored above  $\text{CaH}_2$ . *p*-Cresol or hydroquinone inhibit peroxide formation. The method described by Coetzee and Chang [*Pure Appl Chem* **57** 633 1985] for 1,4-dioxane also applies here. Distillations should always be done in the presence of a reducing agent, e.g.  $\text{FeSO}_4$ . [*Beilstein* **17** H 10, **17** I 5, **17** II 15, **17** III/IV 24, **17**/I V 27.] **It irritates the skin, eyes and mucous membranes, and the vapour should never be inhaled. It is HIGHLY FLAMMABLE, and the necessary precautions should be taken.**

**Rapid purification:** Purification as for diethyl ether.

**l-Tetrahydropalmitine (2,3,9,10-tetramethoxy-6H-dibenzo[a,g]quinolizidine)** [10097-84-4] **M 355.4, m 148-149°, [α]<sub>D</sub><sup>20</sup> -291° (EtOH).** Crystallise it from  $\text{MeOH}$  or  $\text{EtOH}$  by addition of water [see Kametani & Ihara

*J Chem Soc (C)* 530 1967, Bradsher & Dutta *J Org Chem* **26** 2231 1961]. When crystallised from Me<sub>2</sub>CO/Et<sub>2</sub>O, it has **m** 142°. The *hydrate* has **m** 115°(effervescence). The *picrate* has **m** 188°(dec) (from aqueous EtOH). [Bradsher & Datta *J Org Chem* **26** 2231 1961, *Beilstein* **21** II 196, **21** III/IV 2769.]

**Tetrahydropyran (oxane, THP)** [142-68-7] **M 86.1, b 88.0°, d<sub>4</sub><sup>20</sup> 0.885, n<sub>D</sub><sup>20</sup> 1.4202, pK<sup>25</sup> -2.79 (aqueous H<sub>2</sub>SO<sub>4</sub>)**. Dry oxane with CaH<sub>2</sub>, then pass it through a column of silica gel to remove olefinic impurities and fractionally distil it. Free it from peroxides and moisture by refluxing with sodium, then distil it from LiAlH<sub>4</sub>. *Alternatively*, peroxides can be removed by treatment with aqueous ferrous sulfate and sodium bisulfate, followed by solid KOH, and fractional distillation from sodium. [*Beilstein* **17** H 12, **17** I 6, **17** II 18, **17** III/IV 51, **17**/I V 64.]

**Tetrahydro-4H-pyran-4-one** {29943-42-8} **M 100.1, b 57-59°/11mm, 65-66°/15mm, 67-68°/18mm, 73°/20mm, 164.7°/atm, 166-166.5°/atm, d<sub>4</sub><sup>20</sup> 1.0844, n<sub>D</sub><sup>20</sup> 1.4551**. Purify the pyrone by repeated distillation, preferably in a vacuum. [Baker *J Chem Soc* 296 1944, IR: Olsen & Bredoch *Chem Ber* **91** 1589 1958.] The *oxime* has **m** 87-88° and **b** 110-111°/13mm [Cornubert et al. *Bull Soc Chim Fr* 36 1950]. The *4-nitrophenylhydrazone* forms orange-brown needles from EtOH, **m** 186° [Cawley & Plant *J Chem Soc* 1214 1938]. [*Beilstein* **17** I 131, **17** II 287, **17** III/IV 4171, **17**/9 V 21.]

**Tetrahydrothiophene (thiophane)** [110-01-0] **M 88.2, m -96°, b 14.5°/10mm, 40.3°/39.7mm, 120.9°/760mm, d<sub>4</sub><sup>20</sup> 0.997, n<sub>D</sub><sup>20</sup> 1.5289**. The crude material is purified by crystallisation of the mercuric chloride complex to a constant melting point. It is then regenerated, washed, dried, and fractionally distilled. [Whitehead et al. *J Am Chem Soc* **73** 3632 1951.] It has been dried over Na<sub>2</sub>SO<sub>4</sub> and distilled in a vacuum [Roberts & Friend *J Am Chem Soc* **108** 7204 1986]. [*Beilstein* **17** I 5, **17** II 15, **17** III/IV 34, **17**/I V 36.]

**Tetrahydro-4H-thiopyran-4-one** [1072-72-6] **M 116.2, m 60-62°, 61-62°, 64-65°, 65-67°**. Purify it by recrystallisation from diisopropyl ether or petroleum ether and dry it in air. If too impure, then dissolve it in Et<sub>2</sub>O, wash with aqueous NaHCO<sub>3</sub>, then H<sub>2</sub>O, dry (MgSO<sub>4</sub>), filter, evaporate and the residue is recrystallised as before. [Cardwell *J Chem Soc* 715 1949.] The *oxime* can be recrystallised from CHCl<sub>3</sub>/petroleum ether (at -20°) and has **m** 84-85° [Barkenbus et al. *J Org Chem* **20** 871 1955]. The *2,4-dinitrophenylhydrazone* has **m** 186° (from EtOAc) [Barkenbus et al. *J Org Chem* **16** 232 1951]. The *S-dioxide* is recrystallised from AcOH, **m** 173-174° [Fehnel & Carmack *J Am Chem Soc* **70** 1813 1948, *Beilstein* **17** II 287, **17** III/IV 4172, **17**/I V 21].

**Tetramethylene sulfoxide (tetrahydrothiophen 1-oxide)** [1600-44-8] **M 104.2, b 235-237°, d<sub>4</sub><sup>20</sup> 1.175, n<sub>D</sub><sup>20</sup> 1.525**. Shake the oxide with BaO for 4 days, then distil it from CaH<sub>2</sub> under reduced pressure. [*Beilstein* **17** III/IV 36, **17**/I V 38.]

**2,2,6,6-Tetramethylpiperidinyl-1-oxy (TEMPO)** [2564-83-2] **M 156.3, m 36-38°**. Purify TEMPO by sublimation (33°, water aspirator) [Hay & Fincke *J Am Chem Soc* **109** 8012 1987, Keana *Chem Rev* **78** 37 1978].

**2,2,6,6-Tetramethyl-4-piperidone hydrochloride (triacetoneamine)** [33973-59-0] **M 191.7, m 190°(dec), 198-199°(dec), pK<sup>25</sup> 7.90**. Purify the salt by recrystallisation from EtOH/Et<sub>2</sub>O, MeCN or Me<sub>2</sub>CO/MeOH. The *free base* has **m** 37-39° (after sublimation), **b** 102-105°/18mm, and the *hydrate* has **m** 56-58° (wet Et<sub>2</sub>O); the *hydrobromide* has **m** 203° (from EtOH/Et<sub>2</sub>O), and the *picrate* has **m** 196° (from aqueous EtOH). [Sandris & Ourisson *Bull Soc Chim Fr* 345 1958, *Beilstein* **21** H 246, **21** I 273, **21** II 222, **21** III/IV 3278, **21**/6 V 538.]

**1,3,7,9-Tetramethyl uric acid** [2309-49-1] **M 224.2, m 225°, 228°, pK<sub>Est</sub> <0**. Crystallise the uric acid from H<sub>2</sub>O or MeOH. [*Beilstein* **26** H 532, **26** I 156, **26** II 302, **21** III/IV 2623.]

**1,3,5,5-Tetranitrohexahydropyrimidine** [81360-42-1] **M 270.1, m 153-154°**. Crystallise the nitropyrimidine from EtOH (5x) and sublime it (~65°/0.05mm) [Cichra & Adiolph *J Org Chem* **47** 2474 1982, *Labeled Comp Radiopharm* **29** 1197 1991].

**4,7,13,18-Tetraoxa-1,10-diazabicyclo[8.5.5]eicosane (Cryptand 211)** [31250-06-3] M 288.1, b 130°/0.002mm,  $d_4^{20}$  1.097,  $n_D^{20}$  1.505,  $pK_{Est}$  ~7.9. Redistil Cryptand 211, dry it under high vacuum over 24 hours, and store it under nitrogen.

**1,7,10,16-Tetraoxa-4,13-diazacyclooctadecane (4,13-diaza-18-crown-6)** [23978-55-4] M 262.3, m 118-116°,  $pK_{Est}$  ~8.8. Twice recrystallise it from \*benzene/*n*-heptane, and dry it for 24 hours under high vacuum [Weber & Vögtle *Top Curr Chem* (Springer Verlag, Berlin) **98** 1 1981, D'Aprano & Sesta *J Phys Chem* **91** 2415 1987].

**5,10,15,20-Tetraphenylporphyrin (TPP)** [917-23-7] M 614.7, m 450° (sublimes >400°),  $\lambda_{max}$  482nm. Purify TPP by chromatography on neutral (Grade I) alumina, and recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/MeOH or \*C<sub>6</sub>H<sub>6</sub>. It forms complexes with metals. [Yamashita et al. *J Phys Chem* **91** 3055 1987, *Beilstein* **26** III/IV 1958.]

**5,10,15,20-Tetra-4'-pyridinylporphyrin** [16834-13-2] M 618.7, m >300°(dec). Purify it by chromatography on alumina (neutral, Grade I), with CHCl<sub>3</sub>/MeOH (80:20) followed by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/MeOH [Yamashita et al. *J Phys Chem* **91** 3055 1987]. [Kalyanasundaram *Inorg Chem* **23** 2453 1984, Okuno et al. *Synthesis* 537 1980.]

**Tetrathiafulvalene** [31366-25-3] M 204.4, m 122-124°. Recrystallise it from cyclohexane/hexane under an argon atmosphere [Kauzlarich et al. *J Am Chem Soc* **109** 4561 1987]. [*Beilstein* **19/11** V 380.]

**1,2,3,4-(1H)Tetrazole** [288-94-8] M 70.1, m 156°, 157.5-158°,  $pK^{25}$  4.89 (acidic). Crystallise the tetrazole from EtOH and sublime it under high vacuum at ca 120° (care should be taken due to possible **EXPLOSION**). [*Beilstein* **26** H 346, **26** I 108, **26** II 196, **26** III/IV 1652.]

**Thebaine** [115-37-7] M 311.4, m 193°,  $[\alpha]_D^{25}$  -219° (EtOH),  $pK^{20}$  8.15. Crystallise Thebaine from Et<sub>2</sub>O or EtOH. Sublime it at 170-180°. The *hydrochloride* decomposes >182° (from MeOH/Et<sub>2</sub>O). [*Beilstein* **27** II 177, **27** III/IV 2271.] It is a **NARCOTIC**.

**2-Thenoyltrifluoroacetone [1-(2-thienyl)-4,4,4-trifluorobutan-1,3-dione]** [326-91-0] M 222.2, m 42-44°, b 96-98°/9mm,  $pK^{25}$  6.4. Crystallise the dione from hexane or \*benzene. (An aqueous solutions slowly decomposes it). It has  $\nu_{max}$  at 1638(C=O), 1657(C=C)cm<sup>-1</sup>. The *oxime* crystallises from H<sub>2</sub>O or aqueous EtOH. It is used for the determination of Actinides and Lanthanides. [Chaston et al. *Aust J Chem* **18** 673 1956, Jeffrey et al. In *Vogel's Textbook of Quantitative Chemical Analysis* 5<sup>th</sup>edn J Wiley & Sons, p170 1989, *Beilstein* **17** III/IV 5989, **17/11** V 128.]

**2-Thenylamine (2-thiophenemethylamine)** [27757-85-3] M 113.1, b 78.5°/15mm,  $d_4^{20}$  1.137,  $n_D^{20}$  1.5643,  $pK^{30}$  8.92. Distil the amine under reduced pressure (nitrogen), from BaO, through a column packed with glass helices. The *hydrochloride* has m 193-194° (from EtOH/Me<sub>2</sub>CO) and the *picrate* has m 181-182°. [*Beilstein* **18** III/IV 7096.]

**Theobromine (3,7-dimethyl-2,6-dioxopurine)** [83-67-0] M 180.2, m 337° (sublimes slowly at 290° and finally melts at ~351°),  $pK_1^{40}$  -0.16,  $pK_2^{25}$  9.96. It crystallises from H<sub>2</sub>O. Its solubility in H<sub>2</sub>O is 0.06% at 15° and 1.25% at 100°, and it is poorly soluble in organic solvents. It forms salts with heavy metals and is a diuretic, vasodilator and a cardiac stimulant. [Lister *Purines Part II, Fused Pyrimidines* Brown Ed, Wiley-Interscience pp254-225 1971, ISBN 0-471-38205-1, *Beilstein* **26** H 457, **26** I 135, **26** II 264, **26** III/IV 2336.]

**Theophylline (1,3-dimethyl-2,6-dioxopurine, Theocin)** [58-55-9] M 180.2, m 272-274°,  $pK_1^{40}$  -0.24,  $pK_2^{40}$  8.79,  $pK_3$  11.5 (acidic). It crystallises from H<sub>2</sub>O as the *monohydrate* which becomes *anhydrous* above 100°. It is freely soluble in hot H<sub>2</sub>O, but its solubility at 15° is 0.44%. It complexes with heavy metals. It is a diuretic, vasodilator and a cardiac stimulant. [Lister *Purines Part II, Fused Pyrimidines* Brown Ed, Wiley-Interscience pp253-254 1971, ISBN 0-471-38205-1, *Beilstein* **26** H 455, **26** I 134, **26** II 263, **26** III/IV 2331.]

**Thianthrene** [92-85-3] **M 216.3, m 158°**. Crystallise thianthrene from Me<sub>2</sub>CO (charcoal), AcOH or EtOH. It sublimes in a vacuum. [Beilstein 19 H 45, 19 I 619, 19 II 34, 19 III/IV 347, 19/2 V 49.]

**5-Thiazolecarboxaldehyde** [1003-32-3] **M 113.1, b 92-94°/16mm, d<sub>4</sub><sup>20</sup> 1.304, n<sub>D</sub><sup>20</sup> 1.5874, pK<sub>Est</sub> ~0.6**. Dry the aldehyde over Na<sub>2</sub>SO<sub>4</sub> and fractionate it in a vacuum. The 2,4-dinitrophenylhydrazone forms red crystals from MeOH with **m 238-240°**, and the semicarbazone has **m 210-212°** (from MeOH). [Erne et al. *Helv Chim Acta* 34 148 1951, *Beilstein* 27 III/IV 2615.]

**Thiazoline-2-thiol** [96-53-7] **M 119.2, m 106-107°, 106-108°, pK<sub>Est</sub> ~13.0**. Purify the thiol by dissolution in aqueous alkali, precipitation by addition of HCl and then recrystallisation from H<sub>2</sub>O (as needles). [IR: Flett *J Chem Soc* 347 1953 and Mecke et al. *Chem Ber* 90 975 1957, Gabriel & Stelzner *Chem Ber* 28 2931 1895, *Beilstein* 27 III/IV 2540.]

**4-(2-Thiazolylazo)-resorcinol** [2246-46-0] **M 221.2, m 200-202°(dec), 218-219°, λ<sub>max</sub> 500 nm, pK<sub>1</sub><sup>25</sup> 1.25, pK<sub>2</sub><sup>25</sup> 6.53, pK<sub>3</sub><sup>25</sup> 10.76**. Dissolve it in aqueous alkali, extract it with diethyl ether, and re-precipitate it with dilute HCl. The purity is checked by TLC on silica gel using petroleum ether/diethyl ether/EtOH (10:10:1) as the mobile phase. It complexes with Cu<sup>2+</sup> (pH 3-4), Co<sup>2+</sup> and Ni<sup>2+</sup> (pH 7) and Zn<sup>2+</sup>, and Cd<sup>2+</sup> (pH 8.4). [Beilstein 27 III/IV 5988.]

**Thiazolyl blue tetrazolium bromide (MTT, 3-[4,5-dimethyl-2-thiazolyl]-2,5-diphenyl-2H-tetrazolium bromide)** [298-93-1, 2348-71-2] **M 414.3, m 171°**. It is recrystallised by dissolving in MeOH containing a few drops of HBr and then adding dry Et<sub>2</sub>O to complete the crystallisation, wash the needles with Et<sub>2</sub>O and dry them in a vacuum desiccator over KOH. [Beyer & Pyl *Chem Ber* 87 1505 1954, *Beilstein* 27 III/IV 6045.]

**Thietane (trimethylene sulfide)** [287-27-4] **M 74.1, m -64°, -73.2°, b 93.8-94.2°/752mm, 95°/atm, d<sub>4</sub><sup>20</sup> 1.0200, n<sub>D</sub><sup>20</sup> 1.5020**. Purify thietane by preparative gas chromatography on a dinonyl phthalate column. It has also been purified by drying over anhydrous K<sub>2</sub>CO<sub>3</sub>, and distilling through a 25cm glass helices-packed column (for 14g of thietane), then drying over CaSO<sub>4</sub> before sealing it in a vacuum. [Haines et al. *J Phys Chem* 58 270 1954.] It is characterised as the dimethylsulfonium iodide **m 97-98°** [Bennett & Hock *J Chem Soc* 2496 1927]. The *S*-oxide has **b 102°/25mm, n<sub>D</sub><sup>21</sup> 1.5075** [Tamres & Searles *J Am Chem Soc* 81 2100 1959]. [Beilstein 17 I 3, 17 II 12, 17 III/IV 14, 17/1 V 14.]

**2-Thiobarbituric acid** [504-17-6] **M 144.2, m 235°(dec), pK<sub>1</sub><sup>25</sup> 2.25, pK<sub>2</sub><sup>25</sup> 10.72 (2% aqueous EtOH)**. Crystallise it from water. [Beilstein 24 H 476, 24 I 414, 24 II 275, 24 III/IV 1884.]

**1,1'-Thiocarbonyldiimidazole** [6160-65-2] **M 178.1, m 100-102°, 105-106°**. It forms yellow crystals on recrystallisation from tetrahydrofuran or by sublimation at 10<sup>-3</sup> torr (bath temperature 70-80°). It is hydrolysed by H<sub>2</sub>O and should be stored dry. [Staab & Walther *Justus Liebigs Ann Chem* 657 98 1962; Pullukat et al. *Tetrahedron Lett* 1953 1967, Hanessian et al. *Can J Chem* 65 1859 1987, Rajanbabu et al. *J Am Chem Soc* 111 1759 1989.]

**Thiochrome** {2,7-dimethyl-5H-thiachromine-8-ethanol; 3,8-dimethyl-2-hydroxyethyl-5H-thiazolo-[2,3:1',2']pyrimido[4',5'-d]pyrimidine} [92-35-3] **M 262.3, m 227-228°, pK<sub>1</sub><sup>20</sup> 8.11, pK<sub>2</sub><sup>20</sup> 12.6**. Crystallise thiochrome from chloroform. The monohydrochloride has **m 235-236°(dec)** (from EtOH) and the dihydrochloride has **m 237°(dec)**. [Beilstein 27 III/IV 9599.]

**2-Thiocytosine (4-amino-2-mercaptopyrimidine)** [333-49-3] **M 127.2, m 236-237°(dec), 285-290°(dec), pK<sub>1</sub><sup>20</sup> 3.90 (NH<sub>2</sub>), pK<sub>2</sub><sup>20</sup> 11.10 (SH)**. It is recrystallised from hot H<sub>2</sub>O and dried at 100° to constant weight. [Brown *J Appl Chem (London)* 9 203 1959, Russell et al. *J Am Chem Soc* 71 2279 1949.] It is used in transcription and translation studies [Rachwitz & Scheit *Eur J Biochem* 72 191 1977].

**Thioflavine T [2-(4-dimethylaminophenyl)-3,6-dimethylbenzothiazolium chloride]** [2390-54-7] **M 318.9, pK<sub>25</sub> 2.7**. Crystallise the chloride from \*benzene/EtOH (1:1). [Beilstein 27 III/IV 5052.]

**1-Thioflavone (2-phenylthiochromen-4-one)** [784-62-3] **M 238.3, m 129-130°**. This yellow solid is purified by passage through a silica gel column, eluting with \*C<sub>6</sub>H<sub>6</sub>/Me<sub>6</sub>CO, evaporating and crystallising the residue from EtOH. The *sulfoxide* [65373-82-2] has **m 133-135°**, and the *sulfone* [22810-82-2] has **m 136.5-137°** (from EtOH). The *dimethylhydrazone* has **m 111-113°** (from BuOH). It forms easily hydrolysable salts. [Nakazumi et al. *J Heterocycl Chem* **21** 193 1984, Chen et al. *J Org Chem* **51** 3282 1986, Van Allen & Reynolds *J Heterocycl Chem* **8** 807 1971, *Beilstein* **17** I 204, **17** III/IV 5420, **17/10** V 560.]

**6-Thioguanine** [154-42-7] **M 167.2, m >300°, pK<sub>1</sub><sup>23</sup> 8.2 (SH), pK<sub>2</sub><sup>23</sup> 11.6 (acidic, 9-NH)**. It crystallises from H<sub>2</sub>O as needles. It has UV with  $\lambda_{\max}$  at 258 and 347nm (H<sub>2</sub>O, pH 1) and 242, 270 and 322nm (H<sub>2</sub>O, pH 11). [Elion & Hitchings *J Am Chem Soc* **77** 1676 1955, Fox et al. *J Am Chem Soc* **80** 1669 1958.] It is an antineoplastic agent [Kataoka et al. *Cancer Res* **44** 519 1984]. [*Beilstein* **26** III/IV 3926.]

**Thioindigo** [522-75-8] **M 296.2, m >280°**. Adsorb it on silica gel from CCl<sub>4</sub>/\*benzene (3:1), elute with \*benzene, evaporate, crystallise the residue from CHCl<sub>3</sub> and dry it at 60-65° [Wyman & Brode *J Am Chem Soc* **73** 1487 1951; this paper also gives details of purification of other thioindigo dyes]. [*Beilstein* **19** H 137, **19** I 690, **19** II 192, **19** III/IV 2091.]

**Thiomorpholine (tetrahydro-2H-1,4-thiazine)** [123-90-0] **M 103.2, b 110°/100mm, 169°/atm, d<sub>4</sub><sup>20</sup> 1.026, n<sub>D</sub><sup>20</sup> 1.540, pK<sup>25</sup> 9.00**. Purify it by vacuum distillation. The *hydrochloride* has **m 179°** (from isoPrOH or EtOH/Et<sub>2</sub>O/HCl). [Davies *J Chem Soc* 306 1920, *Beilstein* **27** II 4, **2** III/IV 636.]

**Thionine (3,7-diaminophenothiazine, Lauth's violet)** [135-59-1, 581-64-6 (HCl), 78338-22-4 (acetate)] **M 263.7,  $\epsilon_{590}$  6.2 x 10<sup>4</sup> M<sup>-1</sup> cm<sup>-1</sup>, pK<sup>15</sup> 6.9**. The standard biological stain is usually highly pure. It can be crystallised from water or 50% EtOH, then chromatographed on alumina using CHCl<sub>3</sub> as eluent [Shepp et al. *J Phys Chem* **66** 2563 1962]. Dry it overnight at 100° and store it in a vacuum. The *hydrochloride* can be crystallised from 50% EtOH or dilute HCl and aqueous *n*-butanol. Purify it also by column chromatography and washed with CHCl<sub>3</sub> and acetone. Dry it *in vacuo* at room temperature. [*Beilstein* **27** H 391, **27** I 412, **27** II 447, **27** III/IV 5149.]

**Thiooxine hydrochloride (8-mercaptoquinoline hydrochloride)** [34006-16-1] **M 197.7, m 170-175° (dec), pK<sub>1</sub><sup>25</sup> 2.16, pK<sub>2</sub><sup>25</sup> 8.38**. It forms yellow crystals from EtOH. It has pK<sub>a</sub><sup>20</sup> values of 2.05 and 8.29 in H<sub>2</sub>O. It is more stable than thiooxine. [UV: Albert & Barlin *J Chem Soc* 2384 1959]. [*Beilstein* **21** H 99, **21** III/IV 1197, **21/3** V 30.]

**Thiophene** [110-02-1] **M 84.1, f -38.5°, b 84.2°, d<sub>4</sub><sup>20</sup> 1.525, n<sub>D</sub><sup>20</sup> 1.52890, n<sub>D</sub><sup>30</sup> 1.5223**. The simplest purification procedure is to dry thiophene with solid KOH, or reflux it with sodium, and fractionally distil it through a glass-helices-packed column. More extensive treatments include an initial wash with aqueous HCl, then water, drying with CaSO<sub>4</sub> or KOH, and passage through columns of activated silica gel or alumina. Fawcett and Rasmussen [*J Am Chem Soc* **67** 1705 1945] washed thiophene successively with 7M HCl, 4M NaOH, and distilled water, dried with CaCl<sub>2</sub> and fractionally distilled it. \*Benzene was removed by fractional crystallisation by partial freezing, and the thiophene was degassed and sealed in Pyrex flasks. [Also a method is described for recovering the thiophene from the \*benzene-enriched portion.] [*Beilstein* **17** H 29, **17** I 17, **17** II 35, **17** III/IV 234, **17/1** V 297.]

**Thiophene-2-acetic acid** [1918-77-0] **M 142.2, m 63-64°, 76°, b 160°/22mm, pK<sup>25</sup> 3.89, pK<sup>25</sup> 6.43 [MeO(CH<sub>2</sub>)<sub>2</sub>OH-H<sub>2</sub>O/80:20]**. Crystallise the acid from ligroin, hexane and/or distil it in a vacuum. The *amide* has **m 148°** (from H<sub>2</sub>O or petroleum ether). [*Beilstein* **18** H 293, **18** III/IV 4062, **18/6** V 207.]

**Thiophene-3-acetic acid** [6964-21-2] **M 142.2, m 79-80°, pK<sub>Est</sub> ~3.1**. Crystallise the acid from ligroin or H<sub>2</sub>O. [*Beilstein* **18** III/IV 4066.]

**2-Thiophenecarboxaldehyde** [98-03-3] **M 112.2, b 75-77°/11mm, 106°/30mm, 198°/756mm, d<sub>4</sub><sup>20</sup> 1.593, n<sub>D</sub><sup>20</sup> 1.222**. Wash it with 50% HCl and distil it under reduced pressure just before use. It has UV with  $\lambda_{\max}$  234nm (hexane). The *Z-oxime* has **m 144°, 136-138° and 142° (H<sub>2</sub>O)**. [*Beilstein* **17** H 285, **17** I 148, **17** II 313, **17** III/IV 4477, **17/9** V 349.]

**Thiophene-2-carboxylic acid** [527-72-0] **M 128.2, m 129-130°, pK<sup>25</sup> 3.59.** Crystallise the acid from water and dry it in a vacuum. The *amide* has **m 181°(from H<sub>2</sub>O)** and **pK<sup>25</sup> 10.54** (50% aqueous dioxane). [*Beilstein* **18 H 289, 18 I 438, 18 II 269, 18 III/IV 4011, 18/6 V 158.**]

**Thiophene-3-carboxylic acid** [88-31-1] **M 128.1, m 138-139°, pK<sup>25</sup> 6.23(4.11).** Crystallise the acid from water and dry it in a vacuum. [*Beilstein* **18 H 292, 18 III/IV 4053, 18/6 V 199.**] The *amide* has **m 179-180°** (from H<sub>2</sub>O) [*Beilstein* **18 III/IV 4056.**]

**Thiophene-2,5-dicarbonyl dichloride [2,5-bis(chlorocarbonyl)thiophene]** [3857-36-1] **M 209.5, b 102-103°/2mm, 150-152°/11mm, m 43-47°, 45-46°.** Purify it by distillation in a vacuum; or if discoloured, then heat it with SOCl<sub>2</sub> or oxalyl chloride, and distil it in a vacuum. It solidifies on cooling, and can be recrystallised from \*C<sub>6</sub>H<sub>6</sub>/heptane. It provides the *di-phenyl ester* (**m 136-137°**), when treated with phenol at 200° or phenol/pyridine at 100°, followed by adding to cold H<sub>2</sub>O, filtering, drying, and recrystallising from EtOH. [Griffing & Salisbury *J Am Chem Soc* **70** 3416 1968, *Beilstein* **18 H 330, 18 III/IV 4496.**]

**Thiophene-2,5-dicarboxylic acid (2,5-dicarboxythiophene)** [4282-31-4] **M 172.2, m 332-333°(sealed capillary), 358.5-359.5°(corrected, sealed tube), 360°, pK<sub>Est(1)</sub> ~3.3, pK<sub>Est(2)</sub> ~7.4.** It can be precipitated from alkaline solution with acid, filtered and recrystallised from H<sub>2</sub>O, dried and sublimed at 0.0001 mm. Its *mono-methyl ester* crystallises from aqueous MeOH (**m 187-190°, pK<sub>Est(1)</sub> ~3.2**) or petroleum ether (**m 192°**), and sublimes in a vacuum unchanged. With diazomethane/Et<sub>2</sub>O it provides the *di-methyl ester* (**m 148.5-149.5°; m's 146-147° and 152°** were also reported) which crystallises from MeOH or 1:1 aqueous MeOH in flattened needles, and sublimes at 80-95° in high vacuum. This ester has λ<sub>max</sub>(logε) at 275(4.27) in EtOH. [Hartough & Kosak *J Am Chem Soc* **68** 1012 1947, Birkinshaw & Chaplen *Biochem J* **60** 255 1955, Griffing & Salisbury *J Am Chem Soc* **70** 3416 1968, *Beilstein* **18 H 330, 18 III/IV 4496.**]

**Thiopyronine (2,7-dimethylaminothioxanthene chloride hydrochloride)** [2412-14-8] **M 318.9, λ<sub>max</sub> 564nm (ε 78,500) H<sub>2</sub>O, pK<sub>Est</sub> ~ 7.** Purify it as the hydrochloride by recrystallisation from hydrochloric acid forming needles **m 245°** (dec) and UV with λ<sub>max</sub> at 564nm (ε 78,500, H<sub>2</sub>O). [Fanghanel et al. *J Phys Chem* **91** 3700 1987, *Beilstein* **18 H 596, 18 III/IV 7291.**]

**Thiothienoyltrifluoroacetone [1-(2-thienyl)-4,4,4-trifluorobutan-3-one-1-thione]** [4552-64-1] **M 228.2, m 61-62°, 64.5-65°, 73-74°, 74°.** It is easily oxidised and has to be purified before use. This is achieved by recrystallisation from \*benzene or by dissolution in petroleum ether, extraction into 1M NaOH solution, acidification of the aqueous phase with 1-6M HCl solution, back extraction into petroleum ether and final evaporation of the solvent. The purity can be checked by TLC. It is stored in ampoules under nitrogen at 0° in the dark. It crystallises in red crystals from petroleum ether. Its IR has ν<sub>max</sub> at 815m(C-S, C-H), 1260m(C-S), 1570sh(C=C) and 1612s(C=O)cm<sup>-1</sup>. [Müller & Rother *Anal Chim Acta* **66** 49 1973, Chaston et al. *Aust J Chem* **18** 673 1965.]

**2-Thiouracil** [141-90-2] **M 128.2, m 240°(dec), 315°(dec), pK<sub>1</sub><sup>25</sup> 7.75, pK<sub>2</sub><sup>25</sup> 12.7.** Crystallise 2-thiouracil from water or EtOH. [*Beilstein* **24 H 323, 24 I 315, 24 II 171, 24 III/IV 1237.**]

**9H-Thioxanthene-9-one (thioxanthone, thionanthone)** [492-22-8] **M 212.3, m 200-202°, 209°, 212-214°, b 371-373°/712mm.** It forms yellow needles from CHCl<sub>3</sub> or EtOH and sublimes *in vacuo*. It is soluble in CS<sub>2</sub>, hot AcOH, and dissolves in conc H<sub>2</sub>SO<sub>4</sub> to give a yellow colour with green fluorescence in VIS light. The *sulfone* has **m 187°** (from EtOH), and the *hydrazone* has **m 115°** (yellow leaflets from EtOH/\*C<sub>6</sub>H<sub>6</sub>). The *oxime* has **m 194-196°** (from petroleum ether). [Szmant et al. *J Org Chem* **18** 745 1953, Ullmann et al. *Chem Ber* **49** 2509 1916, NMR: Sharpless et al. *Org Magn Res* **6** 115 1974, *Beilstein* **17 H 357, 17 I 191, 17 III/IV 5302, 17/10 V 437.**]

**$\beta$ -Thymidine** [50-89-5] M 242.2, m 185°, 186-188°,  $[\alpha]_D^{20} +19^\circ$  (c 1, H<sub>2</sub>O), pK<sub>2</sub><sup>25</sup> 9.65. Crystallise  $\beta$ -thymidine from ethyl acetate, MeOH/Et<sub>2</sub>O (m 188°) or H<sub>2</sub>O (as 2H<sub>2</sub>O m 189°). It is soluble in water and hot organic solvents. The *picrate* has m 230° (from EtOH). [Beilstein 24 III/IV 1297.]

**Thymine (5-methylpyrimidin-2,4-dione)** [65-71-4] M 126.1, m 326°(dec), pK<sub>1</sub><sup>25</sup> 9.90 (9.82) pK<sub>2</sub><sup>25</sup> >13.0. Crystallise thymine from EtOAc, 10% aqueous EtOH or water. It has m 318-320° after sublimation at 200°/12mm. Purify it by preparative (2mm thick) TLC plates of silica gel, eluting with ethyl acetate/isopropanol/water (75:16:9, v/v; R<sub>F</sub> 0.75). The desired spot is located with a uv lamp, cut the band from the plate, place it in MeOH, shake and filter it through a millipore filter, then evaporate. [Infante et al. *J Chem Soc, Faraday Trans 1* 68 1586 1973, *Beilstein* 24 H 353, 24 I 330, 24 II 183, 24 III/IV 1292.]

**Tinuvin P (2-[2H-benzotriazol-2-yl]-*p*-cresol)** [50936-05-5] M 225.3, m 131-133°, pK<sub>Est(1)</sub>~1.6 (N protonation), pK<sub>Est(2)</sub>~8 (phenolic OH). Recrystallise it from *n*-heptane or Me<sub>2</sub>CO/pentane. [Woessner et al. *J Phys Chem* 81 3629 1985.]

**Toluidine Blue O** [93-31-9] M 305.8, CI 52040,  $\lambda_{\max}$  626nm, pK<sup>25</sup> 7.5. Crystallise the dye from hot water (18ml/g) by adding one and a half volume of alcohol and chilling on ice. Dry it at 100° in an oven for 8-10 hours. [Merrill & Spencer *J Am Chem Soc* 70 3683 1948, *Beilstein* 27 I 417, 27 II 454, 27 III/IV 5161.]

**1,5,7-Triazabicyclo[4.4.0]dec-5-ene (TBD, 1,3,4,6,7,8-hexahydro-2H-pyrimido[1,2-a]pyrimidine)** [5807-14-7] M 139.2, m 125-130°, pK<sup>25</sup> ~16. It crystallises from Et<sub>2</sub>O but readily forms white crystals of the carbonate. It is a strong base (see pK, i.e. about 100 times more basic than tetramethylguanidine). The *picrate* has m 220.5-222° (from EtOH). It forms the 5-nitro derivative m 145-160° that gives a 5-nitro nitrate salt m 100-101° (from EtOH/Et<sub>2</sub>O) and a 5-nitro *picrate* m 144-145° (from H<sub>2</sub>O) [McKay & Kreling *Can J Chem* 35 1438 1957, Schwesinger *Chimia* 39 369 1985, Hilpert et al. *J Chem Soc, Chem Commun* 1401 1983, Kamfen & Eschenmoser *Helv Chim Acta* 72 185 1989]. [*Beilstein* 26 III/IV 60.]

**1,2,4(1H)-Triazole** [288-88-0] M 69.1, m 121°, 260°, pK<sub>1</sub><sup>25</sup> 2.27 (basic), pK<sub>2</sub><sup>25</sup> 10.26 (acidic). Crystallise 1,2,4-triazole from EtOH, H<sub>2</sub>O, EtOAc (m 120.5-121°), or EtOH/\*C<sub>6</sub>H<sub>6</sub>. The *hydrochloride* has m 170°, and the *picrate* has m 163-164° (from H<sub>2</sub>O or CHCl<sub>3</sub>). [Barszcz et al. *J Chem Soc, Dalton Trans* 2025 1986]. [*Beilstein* 26 H 13, 26 II 6, 26 III/IV 35.]

**Tricycloquinazoline** [195-84-6] M 230.3, m 322-323°. Crystallise it repeatedly from toluene, xylene or these solvents mixed with \*C<sub>6</sub>H<sub>6</sub>. It can also be crystallised from CHCl<sub>3</sub> or cymene (m 308-310°), followed by sublimation at 210°/0.15-0.3 Torr in subdued light. [*Beilstein* 26 III/IV 1932.]

**Trifluoperazine dihydrochloride (10-[3-{4-methyl-1-piperazinyl}propyl]-2-trifluoro-methyl-phenothiazine 2HCl)** [440-17-5] M 480.4, m 240-243°, 242-243°, pK<sub>1</sub> 3.9, pK<sub>2</sub> 8.1. Recrystallise the salt from absolute EtOH, filter the crystals, dry them *in vacuo* and store them in tightly stoppered bottles because it is *hygroscopic*. It is soluble in H<sub>2</sub>O but insoluble in \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O and alkaline aqueous solution. It has UV with  $\lambda_{\max}$  at 258 and 307.5nm (log  $\epsilon$  4.50 and 3.50) in EtOH (neutral species). [Craig et al. *J Org Chem* 22 709 1957.] It is a calmodulin inhibitor [Levene & Weiss *J Pharmacol Exptl Ther* 208 454 1978] and is a psychotropic agent [Fowler *Arzneim.-Forsch* 27 866 1977]. [*Beilstein* 27 III/IV 1353.]

**Trigonelline (1-methylnicotinic acid zwitterion)** [535-83-1] M 137.1, m 218°(dec), pK<sub>1</sub><sup>25</sup> 2.10. Crystallise trigonelline (as *monohydrate*) from aqueous EtOH, then dry it at 100°. It also crystallises from H<sub>2</sub>O as the monohydrate with m 230-233°(dec). It has been crystallised from EtOH with m 214-215°(dec). The *hydrochloride* [6138-41-6] M 173.6 has m 258-259°(dec) (from EtOH) [Smisson & Hite *J Am Chem Soc* 81 1201 1959]. The *picrate* crystallises from EtOH with m 204-206°. [Green & Tong *J Am Chem Soc* 78 4896 1956, Kosower & Patton *J Org Chem* 26 1319 1961, *Beilstein* 22 I 504, 22 II 35, 22 III/IV 462, 22/2 V 143.]

**4',5,7-Trihydroxyflavone (apigenin)** [520-36-5] M 270.2, m 296-298°, 300-305°, 345-350° (pK's 7-10, for phenolic OH). Crystallise it from aqueous pyridine or aqueous EtOH. It dyes wool yellow when mixed with Cr ions. [*Beilstein* 18 H 181, 18 I 396, 18 II 178, 18 III/IV 2682, 18/4 V 574.]

**2,2,5-Trimethyl-1,3-dioxane-4,6-dione (Methyl Meldrum's Acid, methylmalonic acid cyclic isopropylidene ester)** [3709-18-0] **M 158.2, m 111-114°, 113-114°, 115°, pK<sup>25</sup> 4.77.** This 'acid' is synthesised and purified in the same way as *Meldrum's acid* [2033-24-1], except that malonic acid is replaced by methylmalonic acid, in 58% yield. [Pihlaja & Seilo *Acta Chim Scand* **22** 3053 1968, Davidson & Bernhard *J Am Chem Soc* **70** 3426 1948]. Its <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has δ at 1.77 (m, isopropylidene Me, *J* = 0.6Hz), 1.85 (m, isopropylidene Me, *J* = 0.6Hz), 1.52 (d, 5-Me, *J* = 7Hz) and 3.81 (q, 5-H, *J* = 7.0Hz) ppm [Schuster & Schuster *Tetrahedron* **25** 199 1969], and the proportion of *enol-form* is apparently not seriously altered by the polarity of the solvent, i.e. MeOH (55.7%), EtOH (1.2%), CDCl<sub>3</sub> (60.8%) and \*C<sub>6</sub>H<sub>6</sub> (58.9%) [Kabachnik et al. *Tetrahedron* **1** 317 1967]. The uncatalysed and acid-catalysed aqueous hydrolysis has been studied in detail with the 5,5-dimethyl-dione (not described here) hydrolysing slightly faster than the 5-methyl-4,6-dione (described here) and the unmethylated acid [Meldrum's acid, 2033-24-1] because, unlike the 5,5-dimethyl-4,6-dione, the latter two can enolise [Pihlaja & Seilo *Acta Chim Scand* **22** 3053 1968.] [*Beilstein* **18** I 480, **19** III/IV 1928, **19/5** V 11.]

**2,2,6-Trimethyl-4H-1,3-dioxin-4-one (diketene acetone adduct)** [5394-63-8] **M 142.2, m 12-13°, b 40°/0.03mm, 65-67°/2mm, 275°/atm, d<sub>4</sub><sup>20</sup> 1.0879, n<sub>D</sub><sup>20</sup> 1.4678.** The reactions of this dioxinone are very similar to those of diketene and it acts as a β-keto-ester synthon. Its purity can be easily assessed by <sup>1</sup>H NMR spectroscopy as it has only three characteristic peaks (see below). Purify it by fractional distillation, preferably under a vacuum. It is a pleasant smelling liquid that is quite stable in the **absence** of alkali. It is slightly soluble in H<sub>2</sub>O that becomes faintly acidic, it gives a red colour with FeCl<sub>3</sub>, and it rapidly reduces alkaline permanganate in alcoholic solution. It is readily prepared by refluxing a mixture of dry acetone (100ml), diketene (100ml, 109g, 1.3moles, see [674-82-8] **toxic**) and *p*-toluenesulfonic acid (0.5g, amount is critical) for 3 hours when the odour of diketene disappears. Excess of acetone is distilled off first followed by the *adduct* (168g, 1.18moles, 91% based on diketene). This procedure should be carried out in an efficient fume cupboard as diketene is **TOXIC**. [Carroll & Bader *J Am Chem Soc* **75** 5400 1953, Naylor (uses ZnCl<sub>2</sub> as catalyst) *J Chem Soc* 244 1945, Dehmlow & Shamout (using quaternary ammonium salts as catalysts) *Justus Liebigs Ann Chem* 1783 1982.] The structure of the adduct has been determined with certainty [Bader et al. *J Org Chem* **21** 821 1956]. Its FT-IR (film) has ν<sub>max</sub> at 1738.7, 1640.0, 1392.8, 1272.6, 1205.3, 1031.5, 901.2, 805.1 and 548.3 cm<sup>-1</sup>; the <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has δ at 1.69 [s, 6H, gem (CH<sub>3</sub>)<sub>2</sub>], 2.00 (s, but d at very high resolution with *J* ~1 Hz, 3H, allyl CH<sub>3</sub>) and 5.21 (s, but q at very high resolution with *J* ~1 Hz, 1H, vinyl H); and the <sup>13</sup>C NMR (CDCl<sub>3</sub>) has δ at 168.65, 161.06, 106.32, 93.81, 25.02 and 19.92. Labelling experiments with (CD<sub>3</sub>)<sub>2</sub>CO showed that isoprenyl acetoacetate may be the key intermediate in the synthesis [Hyatt *J Org Chem* **49** 5102 1984]. In addition to reacting as a diketene reagent, it has been functionalised, e.g. to 6-bromomethyl-2,2-dimethyl-4H-1,3-dioxin-4-one and 6-methylene-4-diethylphosphoryl-2,2-dimethyl-1,3-dioxinane, for the synthesis of natural products such as ikarugamycin and tirandamycin [Boeckman & Thomas *J Org Chem* **47** 2823 1982]. [*Beilstein* **19** IV 1604]

**1',3',3'-Trimethyl-6-nitrospiro[2H-benzopyran-2,2'-indoline]** [1498-88-0] **M 322.4, m 180°.** This photochromic dye crystallises from absolute EtOH [Hinnen et al. *Bull Soc Chim Fr* 2066 1968, Ramesh & Labes *J Am Chem Soc* **109** 3228 1987, Berman et al. *J Am Chem Soc* **81** 5607 1959]. [*Beilstein* **27** III/IV 1460.]

**2,2,4-Trimethyl-6-phenyl-1,2-dihydroquinoline** [3562-69-4] **M 249.3, m 102°.** It is the principal ingredient in Santoflex. Crystallise it three times from absolute EtOH. The 1-phenylcarbamoyl derivative has **m** 148-149.6° (from EtOH). [Hively et al. *Anal Chem* **27** 100 1955, *Beilstein* **20** III/IV 4116.]

**2,4,6-Trimethylpyridine (sym-collidine)** [108-75-8] **M 121.2, m -46°, b 10°/2.7mm, 36-37°/2mm, 60.7°/13mm, 65°/31mm, 170.4°/760mm, 175-178°/atm, d<sub>4</sub><sup>25</sup> 0.9100, n<sub>D</sub><sup>20</sup> 1.4939, 1.4981, n<sub>D</sub><sup>25</sup> 1.4959, pK<sup>25</sup> 6.69(7.45).** Commercial samples may be grossly impure. Likely contaminants include 3,5-dimethylpyridine, 2,3,6-trimethylpyridine and water. Brown, Johnson and Podall [*J Am Chem Soc* **76** 5556 1954] fractionally distilled 2,4,6-trimethylpyridine under reduced pressure through a 40cm Vigreux column and added to 430ml of the distillate slowly, with cooling to 0°, 45g of BF<sub>3</sub>-diethyl etherate. The mixture was again distilled, and an equal volume of dry \*benzene was added to the distillate. Dry HCl was passed into the solution, which was

kept cold in an ice-bath, and the hydrochloride was filtered off. It was recrystallised from absolute EtOH (1.5ml/g) to **m** 286-287° [**m** 256° (sealed tube), also **m** 293-294° subliming slowly]. The free base was regenerated by treatment with aqueous NaOH, then extracted with \*benzene, dried (MgSO<sub>4</sub>) and distilled under reduced pressure. Sisler et al. [*J Am Chem Soc* **75** 446 1953] precipitated trimethylpyridine as its phosphate salt from a solution of the base in MeOH by adding 85% H<sub>3</sub>PO<sub>4</sub>, shaking and cooling. The free base was then regenerated as above. Garrett and Smythe [*J Chem Soc* 763 1903] purified the trimethylpyridine *via* the HgCl<sub>2</sub> complex. It is more soluble in cold than hot H<sub>2</sub>O [the solubility is 20.8% at 6°, 3.5% at 20°, 1.8% at 100°].

*Alternatively*, purify it by dissolving it in CHCl<sub>3</sub>, adding solid K<sub>2</sub>CO<sub>3</sub> and Drierite, filtering and fractionally distilling through an 8in helix-packed column. *The alkali metals, Na, Li or Cs should NOT be used for drying pyridine, and pyridine derivatives, as they form coloured pyridine radical anions leading to bipyridyls.* [see Schmulback et al. *J Am Chem Soc* **90** 6600 1968]. The *sulfate* has **m** 205°, and the *picrate* (from hot H<sub>2</sub>O) has **m** 155-156°. [Frank & Meikle *J Am Chem Soc* **72** 4184 1950, *Beilstein* **20** H 250, **20** I 87, **20** II 164, **20** III/IV 2810, **20/6** V 93.]

**1,3,7-Trimethyluric acid** [5415-44-1] **M 210.2, m 345°(dec), pK<sup>25</sup> 6.0.** Crystallise it from water and dry it at 100° in a vacuum. It has UV with  $\lambda_{\max}$  at 289nm (pH 2.5). [*Beilstein* **26** III/IV 2623.]

**1,3,9-Trimethyluric acid** [7464-93-9] **M 210.2, m 340°(dec), 347°, pK<sup>20</sup> 9.39.** Crystallise it from water and dry it at 100° in a vacuum. [*Beilstein* **26** H 530, **26** II 301, **26** III/IV 2623.]

**1,7,9-Trimethyluric acid** [55441-82-2] **M 210.2, m 316-318°, 345°, pK<sub>Est</sub> ~9.0.** Crystallise the uric acid from water or EtOH, and sublime it *in vacuo*. [*Beilstein* **26** H 530, **26** II 302, **26** III/IV 2623.]

**3,7,9-Trimethyluric acid** [55441-72-0] **M 210.2, m 373-375°(dec), pK<sup>20</sup> 9.39.** Crystallise the uric acid from water and dry it at 100° in a vacuum. It has UV with  $\lambda_{\max}$  at 294nm (pH 2.5). [Bergmann & Dikstein *J Am Chem Soc* **77** 691 1955, *Beilstein* **26** H 530, **26** I 156, **26** II 301, **26** III/IV 2623.]

**1,3,5-Trioxane** [110-88-3] **M 90.1, m 64°, b 114.5°/759mm.** Crystallise 1,3,4-trioxane from sodium-dried diethyl ether or water, and dry it over CaCl<sub>2</sub>. It can also be purified by zone refining. [*Beilstein* **19** H 381, **19** II 392, **19** III/IV 4710, **19/9** V 103.]

**Trioxsalen (2,5,9-trimethyl-7H-furo[3,2-g]benzopyran-7-one)** [3902-71-4] **M 228.3, m 233-235°, 234.5-235°.** Purify trioxsalen by recrystallisation from CHCl<sub>3</sub>. If too impure, it is fractionally crystallised from CHCl<sub>3</sub>/petroleum ether (b 30-60°) using Norit, and finally crystallised from CHCl<sub>3</sub> alone to give colourless prisms, **m** 234.5-235°. It is a photosensitiser so it should be stored in the dark. [UV: Kaufmann *J Org Chem* **26** 117 1961, Baeme et al. *J Chem Soc* 2976 1949, *Beilstein* **19/4** V 472.]

**2,3,5-Triphenyltetrazolium chloride (TTC, TTZ)** [298-96-4] **M 334.8, m 243°(dec).** Crystallise TTZ from EtOH or CHCl<sub>3</sub>, and dry it at 105°. [*Beilstein* **26** H 363, **26** II 216, **26** III/IV 1774.]

**Tripyridyl triazine** [3682-35-7] **M 312.3, m 245-248°, 248-250°.** Purify it by repeated crystallisation from aqueous EtOH. It is a reagent for the determination of Fe(II) and total Fe [Collins et al. *Anal Chem* **31** 1862 1959]. [*Beilstein* **26** III/IV 4192.]

**1,1,1-Tris(acetyloxy)-1,1-dihydro-1,2-benziodoxol-3-(1H)-one (Dess-Martin periodinane, DMP)** [87413-09-0] **M 424.1, m 130-133°, 133-134° (dec).** DMP is essentially the acetylated form of IBX (see above) that renders the latter more soluble in organic solvents. Unlike IBX, it is not explosive on melting or upon hard impact. However, it is subject to hydrolysis in the presence of moisture that may give rise to explosive impurities. It is important that it should be stored in a dry atmosphere and that care should be taken when using it. The original Dess-Martin preparation [Dess-Martin *J Am Chem Soc* **113** 7277 1991] has been improved and is reported here.

1-Hydroxy-1,2-benziodoxol-3(1H)-one (100g, 360mmol, see IBX [61717-82-6]) is added to Ac<sub>2</sub>O (400ml, ~4mol) and TsOH.H<sub>2</sub>O (0.5g), and stirred under N<sub>2</sub> at 80° for 2 hours, then cooled in an ice-water bath. The

white solid that separates in the cold mixture is filtered off onto a fritted glass funnel, rinsed with anhydrous Et<sub>2</sub>O (5 x 50ml), and the crystalline solid (138g, 91%, m 134°) is rapidly transferred to an argon filled amber-glass bottle and stored in a freezer. When exposed to light for several weeks some decomposition occurs so it should be kept away from light. [Ireland & Liu *J Org Chem* **58** 2899 1993.] Its IR (CH<sub>2</sub>Cl<sub>2</sub>) has  $\nu_{\max}$  at 1726.9 (s) and 1707.5 cm<sup>-1</sup>; the <sup>1</sup>H NMR (CDCl<sub>3</sub>, TMS) has  $\delta$  at 2.01 (s, 6H, COCH<sub>3</sub>), 2.33 (s, 3H, COCH<sub>3</sub>), 7.80 (t, 1H,  $J_{\text{HH}} = 7.3$  and 8.5 Hz) and 8.07 (t, 1H,  $J_{\text{HH}} = 7.3$  and 8.5 Hz) (C-4 and C-5), 8.29 (d, 1H,  $J_{\text{HH}} = 8.5$  Hz) and 8.31 (d, 1H,  $J_{\text{HH}} = 8.5$  Hz) (C-3 and C-6); and the <sup>13</sup>C NMR (CDCl<sub>3</sub>) has  $\delta$  at 20.29 (2 COCH<sub>3</sub>), 20.43 (1 COCH<sub>3</sub>), 126.01 (C-2), 126.51, 131.79, 133.81, 135.76, 142.36 (C-1), 166.08 (endocyclic C=O), 173.96 (1 acetate C=O), 175.66 and (2 acetate C=O's). A 0.3M solution of DMP in CH<sub>2</sub>Cl<sub>2</sub> is available commercially.

**DMP** is an extremely useful reagent for the oxidation of primary and secondary alcohols to aldehydes and ketones to completion, without further oxidation to the acids, within ~2 hours at 20°; allylic and benzylic alcohols require ~30 minutes. The reaction is increased rapidly on addition of an alcohol. Furan rings or sulfides and vinyl ethers are unreactive. Geraniol is oxidised to geranial without rearrangement to nerol, and it oxidises *N*-benzylbenzamide to benzaldehyde. The reagent is usually added in an appropriate anhydrous solvent (e.g. CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub>, MeCN, etc.) followed by the substrate. Workup can be by adding Et<sub>2</sub>O followed by aqueous NaOH which decomposes the reagent to iodobenzoate, or for base-sensitive substrates NaHCO<sub>3</sub> and sodium thiosulfate (Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub>) are added, which also remove iodine. A pyridine buffer can be used to keep near-neutral conditions throughout the entire oxidation followed by a thiosulfate work up procedure. [see Dess-Martin *J Am Chem Soc* **113** 7277 1991, and papers cited in this reference.] In the peptide field, substances like *N*-Fmoc phenylglycinol are oxidised to *N*-Fmoc phenylglycinal almost quantitatively in water-saturated CH<sub>2</sub>Cl<sub>2</sub> with this reagent [Myers et al. *Tetrahedron Lett* **41** 1359 2000].

**1,3,5-Trithiane (trithioformaldehyde)** [291-21-4] **M 138.3, m 216-218°(dec)**. Crystallise it from AcOH or toluene, after Soxhlet extraction with toluene (30g/300ml) [*Beilstein* **19** III/IV 4711, **19/9** V 105.]

**3-Tropanol (Tropine)** [120-29-6] **M 141.2, m 63°, 64-66°, b 229°/760mm, pK<sup>15</sup> 3.80**. Distil 3-tropanol in steam and crystallise it from Et<sub>2</sub>O or toluene/petroleum ether. It is soluble in H<sub>2</sub>O, EtOH, CHCl<sub>3</sub>, and sublimes at 60°/0.1mm. *Hygroscopic*. A 0.05M solution in H<sub>2</sub>O has a pH of 11.5. The *hydrochloride* has **m 280°** (from EtOH/Et<sub>2</sub>O). [*Beilstein* **21** H 16, **21** I 197, **21** II 17, **21** III/IV 168, **21/1** V 219.]

**Tryptamine [3-(2-aminoethyl)indole]** [61-54-1] **M 160.1, m 116°, pK<sub>1</sub><sup>25</sup> -6.31 (aqueous H<sub>2</sub>SO<sub>4</sub>, diprotonation), pK<sub>Est(2)</sub>~4.9, pK<sub>3</sub><sup>25</sup> 16.60 (acidic indole NH)**. Crystallise tryptamine from \*benzene, Et<sub>2</sub>O (m 114°) or petroleum ether (m 118°). It has UV with  $\lambda_{\max}$  at 222, 276, 282 and 291nm (EtOH) and 226, 275, 281 and 290nm (HCl). [*Beilstein* **22** II 346, **22** III/IV 4319, **22/10** V 45.]

**Tryptamine hydrochloride** [343-94-2] **M 196.7, m 252-253°**. Crystallise the salt from EtOH/water or EtOH/Et<sub>2</sub>O. See previous entry for UV. [*Beilstein* **22** II 347, **22** III/IV 4319, **22/10** V 46.]

**Tryptophol [3-(2-hydroxyethyl)indole]** [526-55-6] **M 161.2, m 59°, b 174°/2mm**. Crystallise it from diethyl ether/petroleum ether, \*C<sub>6</sub>H<sub>6</sub>, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. The *picrate* has **m 100-101°** (from \*C<sub>6</sub>H<sub>6</sub>). [*Beilstein* **21** I 218, **21** II 49, **21** III/IV 788, **21/3** V 61.]

**(+)-Tubocurarine chloride (5H<sub>2</sub>O)** [57-94-3] **M 771.7, m 274-275°(dec) (anhydrous), [α]<sub>D</sub><sup>20</sup> +235° (c 0.5, H<sub>2</sub>O), pK<sub>Est(1)</sub>~8.5, pK<sub>Est(2)</sub>~8.8**. Crystallise this chloride from water. It forms various hydrates. The *hydrochloride pentahydrate* has **m 268-269°** (from H<sub>2</sub>O) and [α]<sub>D</sub><sup>21</sup> +190° (0.5, H<sub>2</sub>O). Its solubility in H<sub>2</sub>O at 25° is 50mg/ml. [*Beilstein* **27** II 897, **27** III/IV 8727.]

**Umbelliferone (7-hydroxycoumarin)** [93-35-6] **M 162.2, m 225-228°, 230-233°, pK<sub>Est</sub> ~8.0**. It crystallises from water (m 232-232.2°) or EtOH (m 232°). It sublimes at 160°/0.001mm. Fluorescence: Em<sub>max</sub> 452nm/Exc<sub>max</sub> 325nm in 50% EtOH. [*Beilstein* **18** H 27, **18** I 306, **18** II 16, **18** III/IV 294, **18/1** V 386.]

**Uracil (pyrimidine-2,4(1*H*)-dione)** [66-22-8] **M 122.1, m 335°(dec), pK<sub>1</sub><sup>25</sup> 9.43, pK<sub>2</sub><sup>25</sup> 13.3-14.2.** Uracil crystallises from water (**m** 339-341°) and **m** 338° after sublimation in high vacuum. Its solubility in H<sub>2</sub>O at 20° is 1g/300ml. [*Beilstein* 24 H 312, 24 I 312, 24 II 169, 24 III/IV 1193.]

**Uramil (5-aminobarbituric acid)** [118-78-5] **M 143.1, m 310-312°, 320°, >400°(dec), pK<sub>Est(1)</sub>~3.9, pK<sub>Est(2)</sub>~8.0, pK<sub>Est(3)</sub>~12.5.** It crystallises from water. It has also been purified by dissolving it in aqueous ammonia and precipitating it by dropwise addition of formic acid. The solid is collected and dried in a vacuum at 100°. [Hartman & Sheppard *Org Synth Coll Vol II* 617 1943, *Beilstein* 25 H 492, 25 I 704, 25 III/IV 4228.]

**Uric acid** [69-93-2] **M 168.1, m >300° (dec) pK<sub>1</sub> 5.75, pK<sub>2</sub> 10.3.** Crystallise uric acid from hot distilled H<sub>2</sub>O (the solubility in H<sub>2</sub>O is 1part/39,000parts at 18° and 1part/2,000parts at 100°). It is best purified by dissolving in an alkaline solution and acidifying with dilute HCl and drying it at 100° in a vacuum. [Bergmann & Dikstein *J Am Chem Soc* 77 691 1955, Lister *Purines Part II, Fused Pyrimidines* Brown Ed, Wiley-Interscience pp256-257 1971, ISBN 0-471-38205-1, *Beilstein* 26 H 513, 26 I 151, 26 II 293, 26 III/IV 2619.]

**β-Uridine** [58-96-8] **M 244.2, m 165°, [α]<sub>D</sub><sup>20</sup> +10.0° (c 1.6, H<sub>2</sub>O), pK<sub>25</sub> 9.51 (9.25).** Crystallise β-uridine from aqueous 75% MeOH or EtOH (**m** 165-166°). [*Beilstein* 24 III/IV 1202.]

**Urocanic acid (4-imidazolylacrylic acid)** [104-98-3] **M 138.1, m 225°, 226-228°, pK<sub>Est(1)</sub>~2.5, pK<sub>Est(2)</sub>~6, pK<sub>Est(3)</sub>~11.** Crystallise the acid from water and dry it at 100°. The *trans-isomer* [3465-72-3] has **m** 225° (229-230°, 230-231° or 231°(dec, from H<sub>2</sub>O) and pK<sub>1</sub> 3.5 and pK<sub>2</sub> 5.6, and the *picrate* has **m** 225°(dec, from H<sub>2</sub>O). The *cis-isomer* [7699-35-6] has **m** 175-176° (178-179° or 180-184° dec, from H<sub>2</sub>O) and pK<sub>1</sub> 3.0 and pK<sub>2</sub> 6.7, and the *picrate* has **m** 204° (from H<sub>2</sub>O). [*Beilstein* 25 H 124, 25 I 536, 25 II 121, 25 III/IV 786.]

**δ-Valerolactam (2-piperidone)** [675-20-7] **M 99.1, m 38.5-39.5°, 39-40°, 40°, b 81-82°/0.1mm, 136-137°/15mm, pK<sub>25</sub> 0.75 (in AcOH).** Purify it by repeated fractional distillation. [Cowley *J Org Chem* 23 1330 1958, Reppe et al. *Justus Liebigs Ann Chem* 596 198 1955, IR: Huisgen et al. *Chem Ber* 90 1437 1957.] The *hydrochloride* [5174-67-4] has **m** 183-184° (from isoPrOH or EtOH/Et<sub>2</sub>O) [Hurd et al. *J Org Chem* 17 865 1952], and the *oxime* has **m** 122.5° (from petroleum ether) [Behringer & Meier *Justus Liebigs Ann Chem* 607 67 1957]. The *N-benzoyl* derivative [4252-56-6] has **m** 115° (from CHCl<sub>3</sub>/petroleum ether), the *N-methyl* derivative [931-20-4] is a water soluble hygroscopic liquid with **b** 115°/14mm, and the *N-methyl hydrochloride* [87243-73-0] crystallises from from EtOH/Et<sub>2</sub>O with **m** 115°. The *picrate* has **m** 92-93°. [*Beilstein* 21 H 239, 21 III/IV 3170, 21/6 V 396.]

**δ-Valerolactone (tetrahydro-2*H*-pyran-2-one)** [542-28-9] **M 100.1, m -13°, -12°, b 88°/4mm, 97°/10mm, 124°/24mm, 145-146°/40mm, 229-229.5°/atm, d<sub>4</sub><sup>20</sup> 1.1081, n<sub>D</sub><sup>20</sup> 1.4568.** Purify the δ-lactone by repeated fractional distillation. Its IR has **v<sub>max</sub>** at 1750 (in CS<sub>2</sub>), 1732 (in CHCl<sub>3</sub>), 1748 (in CCl<sub>4</sub>) and 1733 (in MeOH) cm<sup>-1</sup> [Huisgen & Ott *Tetrahedron* 6 253 1959, Linstead & Rydon *J Chem Soc* 580 1933, Jones et al. *Can J Chem* 37 2007 1959]. [*Beilstein* 17 H 235, 17 II 287, 17 III/IV 4169, 17/9 V 17.]

**γ-Valerolactone (± 4,5-dihydro-5-methyl-2(3*H*)-furanone)** [108-29-2] **M 100.1, m -37°, 36°, b 82-85°/10mm, 102-103°/28mm, 125.3°/68mm, 136°/100mm, 205.75-206.25°/ 754mm, d<sub>4</sub><sup>20</sup> 1.072, n<sub>D</sub><sup>20</sup> 1.4322.** Purify the γ-lactone by repeated fractional distillation [Boorman & Linstead *J Chem Soc* 577, 580 1933]. Its IR has **v<sub>max</sub>** at 1790 (CS<sub>2</sub>), 1775 (CHCl<sub>3</sub>) cm<sup>-1</sup> [Jones et al. *Can J Chem* 37 2007 1959]. The BF<sub>3</sub>-complex distils at 110-111°/20mm [Reppe et al. *Justus Liebigs Ann Chem* 596 179 1955]. It is characterised by conversion to γ-hydroxy-*n*-valeramide on treatment with NH<sub>3</sub>, which has **m** 51.5-52° (by slow evaporation of a CHCl<sub>3</sub> solution). [*Beilstein* 17 H 235, 17 I 131, 17 II 288, 17 III/IV 4176, 17/9 V 24.]

**(±)-Vinclozolin [3-(3,5-dichlorophenyl)-5-methyl-5-vinylloxazolidine-2,4-dione]** [50471-44-8] **M 286.1, m 108°.** Crystallise the fungicide from Me<sub>2</sub>CO/H<sub>2</sub>O. Its solubility at 20° (w/w%) is 44 (Me<sub>2</sub>CO), 32 (CHCl<sub>3</sub>), 25 (EtOAc) and 10 (H<sub>2</sub>O). It irritates the eyes and skin. [GP 2,207,576 1973, *Chem Abstr* 79 137120 1973.]

**N-Vinylcaprolactam** [2235-00-9] **M 139.2, m 35-38°(polym), b 95-95.5°/4mm, 128°/21mm,  $d_4^{20}$  1.0287,  $n_D^{20}$  1.5133.** Distil it under vacuum and with 0.0015% of 4-*tert*-butylcatechol as stabilizer. [Beilstein 21 III/IV 3207.]

**N-Vinylcarbazole** [1484-13-5] **M 193.3, m 66°.** Crystallise *N*-vinylcarbazole repeatedly from MeOH in amber glassware. It sublimes in a vacuum. [Beilstein 20 II 282, 20 III/IV 3830, 20/8 V 19.]

**Vinylene carbonate (1,3-dioxol-2-one)** [872-36-6] **M 86.1, m 22°, b 76-78°/37mm, 165°/~760mm.** Purify it by zone melting, or distillation, and stabilise it with 0.5% of 2,6-di-*tert*-butyl-*p*-cresol. [Beilstein 19 III/IV 1597, 19/4 V 72.]

**2-Vinylpyridine monomer** [100-69-6] **M 105.1, b 79-82°/29mm, d 0.974, n 1.550,  $pK^{25}$  4.92(4.98).** Steam distil it, then dry it with  $MgSO_4$  and distil it in a vacuum. [Beilstein 20 H 256, 20 III/IV 2884, 20/6 V 211.]

**4-Vinylpyridine monomer** [100-43-6] **M 105.1, b 40-41°/1.4mm, 54°/5mm, 58-61°/12mm, 68°/18mm, 79°/33mm,  $d_4^{20}$  0.9836,  $n_D^{20}$  1.5486,  $pK^{25}$  5.62.** Purify the monomer by fractional distillation under a good vacuum and in a  $N_2$  atmosphere; store it in sealed ampoules under  $N_2$ , and keep it in the dark at -20°. The *picrate* has **m 175-176°.** [UV: Coleman & Fuoss *J Am Chem Soc* 77 5472 1955, Overberger et al. *J Polymer Sci* 27 381 1958, Petro & Smyth *J Am Chem Soc* 79 6142 1957.] It is used for alkylating SH groups in peptides [Anderson & Friedman *Can J Biochem* 49 1042 1971, Cawins & Friedman *Anal Biochem* 35 489 1970]. [Beilstein 20 II 170, 20 III/IV 2887, 20/6 V 213.]

**Viologen (4,4'-dipyridyl dihydrochloride)** [27926-72-3] **M 229.1, m 278° (also reported m 302-306°, >300°, with sublimation),  $pK_1^{20}$  3.17,  $pK_2^{20}$  4.82.** Purify viologen by precipitation on adding excess of acetone to a concentrated solution of it in aqueous MeOH. It has also been recrystallised several times from MeOH or *iso*-propanol and dried at 70° under vacuum for 24 hours [Prasad et al. *J Am Chem Soc* 108 5135 1986], and recrystallised three times from MeOH/*isopropanol* [Stramel & Thomas *J Chem Soc, Faraday Trans* 82 799 1986, Michaelis & Hill *J Am Chem Soc* 55 1481 1933, Tilford et al. *J Am Chem Soc* 70 4005 1948]. [Beilstein 23 I 49.]

**Visnagin (4-methoxy-7-methyl-5H-furo[3,2-g][1]benzopyran-5-one)** [82-57-5] **M 230.2, m 142-145°.** Crystallise visnagin from water. It is soluble in  $CHCl_3$  but slightly soluble in EtOH. [Aneja et al. *Tetrahedron* 3 230 1958, Beilstein 19 III/IV 2640.]

**9H-Xanthene (dibenzopyran)** [92-83-1] **M 182.2, m 100.5°, 101-102.5°, b 310-312°/760mm.** Crystallise dibenzopyran from \*benzene, MeOH or EtOH. [Beilstein 17 H 73, 17 I 30, 17 II 72, 17 III/IV 614, 17/2 V 252.]

**Xanthine (2,6-dihydroxypurine, purine-2,6(1H,3H)dione)** [69-89-6] **M 152.1,  $pK_1$  0.8 [protonation of imidazole 7(9)NH],  $pK_2$  7.44 [monoanion 1(3)NH],  $pK_3$  11.12 [dianion 1,3-N<sup>2-</sup>].** The *monohydrate* separates in a microcrystalline form on slow acidification with acetic acid of a solution of xanthine in dilute NaOH. It is also precipitated by addition of conc  $NH_3$  to its solution in hot 2N HCl (charcoal). After washing with  $H_2O$  and EtOH, it is dehydrated by heating above 125°. Its solubility in  $H_2O$  is 1 in 14,000 parts at 16° and 1 in 1,500 parts of boiling  $H_2O$ , and separates as plates. It has no **m**, but the *perchlorate* has **m 262-264°** [Lister *Purines Part II, Fused Pyrimidines* Brown Ed, Wiley-Interscience pp 252-253 1971, ISBN 0-471-38205-1]. [Beilstein 26 H 447, 26 I 131, 26 II 260, 26 III/IV 2327.]

**9-Xanthone (9-xanthenone)** [90-47-1] **M 196.2, m 175.6-175.4°.** Crystallise xanthone from EtOH (25ml/g) and dry it at 100°. It has also been recrystallised from *n*-hexane three times and sublimed *in vacuo*. [Saltiel *J Am Chem Soc* 108 2674 1986]. [Beilstein 17 H 354, 17 I 190, 17 II 378, 17 III/IV 5292, 17/10 V 430.]

**Xanthosine (2H<sub>2</sub>O) [9-(β-D-ribose)purin-2,6(1H,3H)-dione]** [5968-90-1] **M 320.3,  $[\alpha]_D^{20}$  -53° (c 8, 0.3M NaOH),  $pK_1^{25}$  <2.5,  $pK_2^{25}$  5.67,  $pK_3^{25}$  12.85.** It crystallises from EtOH (anhydrous) or water (as dihydrate).

[Howard et al. *J Chem Soc* 232 1949, *Beilstein* 31 H 28, 26 III/IV 2428.]

**Xanthurenic acid (5,8-dihydroxyquinoline-2-carboxylic acid)** [59-00-7] **M 205.2, m 286<sup>o</sup>, 290-295<sup>o</sup>(dec), 297-29<sup>o</sup>(dec), pK<sub>Est(1)</sub>~ 1.5, pK<sub>Est(2)</sub>~ 4.9, pK<sub>Est(3)</sub>~ 9.8.** It is precipitated by the addition of 2N formic acid to its solution in hot 2M ammonia (charcoal). The solid is filtered off, dried in a vacuum at ~80<sup>o</sup> in the dark. Its UV (H<sub>2</sub>O) has  $\lambda_{\max}$  nm ( $\epsilon$  M<sup>-1</sup>cm<sup>-1</sup>) at 243 (30,000) and 342 (6,500). The *methyl ester* has **m 262<sup>o</sup>** (from MeOH). It forms Cu<sup>2+</sup>, Zn<sup>2+</sup>, Fe<sup>2+</sup> and Fe<sup>3+</sup> salts. [*Beilstein* 22 III/IV 2513.]

**Xanthydrol** [90-46-0] **M 198.2, m 122-123<sup>o</sup>, 123-124<sup>o</sup>, 124-126<sup>o</sup>.** Crystallise xanthydrol from EtOH and dry it at 40-50<sup>o</sup>. [*Beilstein* 17 H 129, 17 I 72, 17 II 146, 17 III/IV 1602, 17/4 V 502.]

**Xylenol Orange {3H-2,1-benzoxathiol-3-ylidene-bis-[(6-hydroxy-5-methyl-*m*-phenylene)-methylnitrilo]-tetraacetic acid, S,S-dioxide}** [1611-35-4] **M 672.6, m 210<sup>o</sup>(dec),  $\epsilon_{578}$  6.09 x 10<sup>4</sup> (pH 14),  $\epsilon_{435}$  2.62 x 10<sup>4</sup> (pH 3.1), pK<sub>1</sub> -1.74, pK<sub>2</sub> -1.09 (aqueous H<sub>2</sub>SO<sub>4</sub>-HNO<sub>3</sub>), pK<sub>3</sub> 2.58, pK<sub>4</sub> 3.23, pK<sub>5</sub> 6.46, pK<sub>6</sub> 10.46, pK<sub>7</sub> 12.28.** It is generally contaminated with starting material (cresol red) and semi-xylenol orange. Purify it by ion-exchange chromatography using DEAE-cellulose, eluting with 0.1M NaCl solution, which will give the sodium salt [3618-43-7]. Cresol Red, semi-xylenol orange and iminodiacetic acid bands elute first. This procedure will give the sodium salt of the dye.

To obtain the *free acid*, dissolve the salt in H<sub>2</sub>O and acidify it with AcOH. Filter it off, wash it with H<sub>2</sub>O and dry it first in air and then in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> in the dark [Sato et al. *Anal Chim Acta* 94 317 1977]. [*Beilstein* 19 II 111, 19 III/IV 1135, 19/3 V 461.]

**$\alpha$ -Yohimbine (17 $\alpha$ -Hydroxy-20 $\alpha$ -yohimban-16 $\alpha$ -carboxylic acid methyl ester, Rauwolscine)** [146-48-5] **M 354.5, m 234<sup>o</sup>, 235-237<sup>o</sup>(dec), 278<sup>o</sup>(dec),  $[\alpha]_D^{20}$  +55.6<sup>o</sup> (c 2, EtOH), +108<sup>o</sup> (pyridine), pK<sub>1</sub><sup>22</sup> 3.0, pK<sub>2</sub><sup>22</sup> 7.45.** Crystallise the *Rauwolfia* alkaloid from EtOH or dilute EtOH (needles), and dry it in a vacuum to remove EtOH of crystallisation. It is soluble in MeOH, CHCl<sub>3</sub>, warm \*C<sub>6</sub>H<sub>6</sub>, moderately soluble in Et<sub>2</sub>O but difficultly soluble in H<sub>2</sub>O. Its has UV (MeOH) with  $\lambda_{\max}$  nm(log $\epsilon$ ) at 226(4.56), 280(3.88) and 291(380). The *hydrochloride* [45-19-0] crystallises in plates or prisms from EtOH with **m 302<sup>o</sup>(dec),  $[\alpha]_D^{20}$  +105<sup>o</sup> (H<sub>2</sub>O).** Its solubility in H<sub>2</sub>O is 83% (giving a neutral solution), and 0.25% in EtOH. [Van Tamelen et al. *J Am Chem Soc* 91 7315 1969, Stork & Guthikonda *J Am Chem Soc* 94 5109 1972, Wenkert et al. *J Am Chem Soc* 100 4894 1978 and *J Am Chem Soc* 101 5370 1982, *Beilstein* 25 II 201, 25 III/IV 1234.], It is pharmacologically active being a 5-HT<sub>1A</sub> serotonin receptor agonist and an  $\alpha_2$ -adrenoceptor antagonist.

For  $\delta$ -yohimbine and  $\gamma$ -yohimbine see ajmalicine [483-04-5] and ajmaline [4360-12-7] above.

## MISCELLANEOUS As, B, P, Si, S, Se and Te COMPOUNDS.

This section contains miscellaneous organic compounds of As, B, P, Si, S, Se, Te, and ammonium and metal salts of their acids, and salts of their bases. See other sections and chapters for further entries of sulfur, phosphorous and silicon compounds; and for sulfur heterocyclic compounds see section on "Heterocyclic Compounds",

**Acetyltriphenylphosphonium chloride** [1235-21-8] **M 354.8, m 237-238°, 244-246°(dec)**. Recrystallise it from  $\text{CHCl}_3$ / $\text{C}_6\text{H}_6$ /petroleum ether (b 60-80°) or by dissolving it in  $\text{CHCl}_3$  and pouring it into dry  $\text{Et}_2\text{O}$ . It has UV(EtOH) with  $\lambda_{\text{max}}$  nm( $\epsilon$ ) at 255(3,600), 262(3,700), 268(4,000) and 275(3,100). The *iodide salt* crystallises from  $\text{H}_2\text{O}$  and has **m** 207-209°. [Ramirez & Dershowitz *J Org Chem* **22** 41 1957.] It is an **IRRITANT** and is *hygroscopic*. When shaken with a 10% aqueous solution of  $\text{Na}_2\text{CO}_3$  (8 hours) it gives *acetylmethylene triphenyl phosphorane* which is recrystallised from  $\text{MeOH}/\text{H}_2\text{O}$ , and after drying at 70°/0.1mm has **m** 205-206°. It has UV with  $\lambda_{\text{max}}$  nm( $\epsilon$ ) at 268 (6600), 275 (6500) and 288 (5700), and its IR has  $\nu_{\text{max}}$  at 1529(s), 1470(m), 1425(s), 1374(m), 1105(s) and 978(s) ( $\text{cm}^{-1}$ ). [Ramirez & Dershowitz *J Org Chem* **22** 41, 44 1957, *Beilstein* **16** H 761, **16** II 373.]

**3R,4R,1'R-4-Acetoxy-3-[1-(tert-butylmethylsilyloxy)ethyl]-2-azetinone** [76855-69-1] **M 287.4, m 107-108°,  $[\alpha]_{\text{D}}^{20} +55^\circ$  (c 0.5, toluene),  $[\alpha]_{\text{D}}^{20} +53.7^\circ$  (c 1.04,  $\text{CHCl}_3$ )**. Purify it by chromatography on silica gel (3 x 14cm) for 50g of ester using 20% EtOAc in *n*-hexane. The eluate is evaporated, and the residue is recrystallised from hexane (white fluffy crystals). [Leanza et al. *Tetrahedron* **39** 2505 1983.]

**N-Acetyl-4-hydroxy-*m*-arsanilic acid (Acetarsol, 3-acetamido-4-hydroxyphenylarsonic acid)** [97-44-9] **M 275.1, m 240-250°,  $\text{pK}_1$  3.73,  $\text{pK}_2$  7.9,  $\text{pK}_3$  9.3**. It crystallises from water in colourless prisms. It decomposes slowly on prolonged boiling in  $\text{H}_2\text{O}$  or dilute alkalis. The *N*-propionyl derivative recrystallises from  $\text{H}_2\text{O}$  with **m** 228-229°(dec). [Raiziss & Fisher *J Am Chem Soc* **48** 1323 1926, Hewitt & King *J Chem Soc* 823 1926, *Beilstein* **16** I 491, **16** II 521, **16** III 1129.]

**Alizarin Red S (3,4-dihydroxy-9,10-dioxo-2-anthracene sulfonic acid, Na salt.  $\text{H}_2\text{O}$ )** [130-22-3] **M 360.4,  $\text{pK}_1^{25} <1$ ,  $\text{pK}_2^{25} 5.49$ ,  $\text{pK}_3^{25} 10.85$  (11.01)**. Commercial samples contain large amounts of sodium and potassium chlorides and sulfates. It is purified by passing through a Sephadex G-10 column (size exclusion column), followed by elution with water, then 50% aqueous EtOH [King & Pruden *Analyst (London)* **93** 601 1968]. Finally dissolve it in EtOH and precipitate it with  $\text{Et}_2\text{O}$  several times [Sacconi *J Phys Chem* **54** 829 1950, polarography: Furnam & Stone *J Am Chem Soc* **70** 3055 1948, *Beilstein* **11** IV 682.]

**Allyl trimethylsilane (2-propenyltrimethylsilane)** [762-72-1] **M 114.3, b 83.0-84.5°, 84-88°, 85.5-86.0°,  $d_4^{20} 0.713$ ,  $n_{\text{D}}^{20} 1.405$** . Fractionate it through an efficient column at atmospheric pressure. If impure, dissolve it in THF, shake it with  $\text{H}_2\text{O}$  (2x), dry ( $\text{Na}_2\text{SO}_4$ ), filter and fractionate it. [Cudlin & Chvalovský *Col Czech Chem Commun* **27** 1658 1962, *Beilstein* **4** IV 3927.]

**Ammonium dodecylsulfate (ammonium laurylsulfate)** [2235-54-3] **M 283.4**. Recrystallise it first from 90% EtOH and then twice from absolute EtOH, and finally dry it in a vacuum. [*Beilstein* **1** III 1786.]

**Ammonium tetraphenylborate** [14637-34-4] **M 337.3, m ca 220°(dec)**. Dissolve it in aqueous  $\text{Me}_2\text{CO}$  and allow crystallisation to proceed slowly; otherwise very small crystals are formed. No trace of  $\text{Me}_2\text{CO}$  is left in the crystals after drying at 120° [Davies & Staveley *Trans Faraday Soc* **53** 19 1957]. Also, the salt can be precipitated from a dilute AcOH solution of sodium tetraphenylborane in the presence of  $\text{NH}_4^+$  ions. After standing for 5 minutes, the precipitate is filtered off onto a sintered porcelain crucible, washed with very dilute AcOH and dried at room temperature for at least 24 hours [Vendlandt *Anal Chem* **28** 1001 1956]. Alternatively, a solution of sodium tetraphenylborane (5% excess) in  $\text{H}_2\text{O}$  is added to  $\text{NH}_4\text{Cl}$  solution. After 5 minutes the precipitate is collected, washed several times with  $\text{H}_2\text{O}$  and recrystallised from aqueous  $\text{Me}_2\text{CO}$ . [Howick & Pflaum *Analyt Chim Acta* **19** 342 1958, *Beilstein* **16** IV 1625.]

**9-Anthraceneboronic acid** [100622-34-2] M 222.0, m 203-250°. Crystallise the boronic acid from dilute HCl (m 180-184°). The *disodium salt* has m 209-213°. [Beilstein 16 IV 1679.]

**Anthraquinone Blue B (Acid Blue 45, 1,5-diamino-4,8-dihydroxy-9,10-anthraquinone-3,7-disulfonic acid di-Na salt)** [2861-02-1] M 474.3, m >300°, CI 63010,  $\lambda_{\max}$  595nm,  $\text{pK}_{\text{Est}(1)} \sim <0$ ,  $\text{pK}_{\text{Est}(2)} \sim 2$ ,  $\text{pK}_{\text{Est}(3)} \sim 9$ . Purify it by salting out an aqueous solution three times with sodium acetate, followed by repeated extraction with EtOH [McGrew & Schneider *J Am Chem Soc* 72 2547 1950, Beilstein 14 H 706, 725].

**Anthraquinone Blue RXO** [4403-89-8] M 445.5. Purify the dye by salting out an aqueous solution three times with sodium acetate, followed by repeated extraction with EtOH [McGrew & Schneider *J Am Chem Soc* 72 2547 1950]. [Beilstein 14 H 706, 725.]

**Anthraquinone Green G [Acid Green 25, Alizarin Cyanine Green F, 1,4-bis-(4-methyl-2-sulfophenyl-1-amino)-9,10-anthraquinone di-Na salt]** [4403-90-1] M 624.6, m 235-238°, CI 61570,  $\lambda_{\max}$  642nm,  $\text{pK}^{25} >0$ . Purify it by salting out three times from an aqueous solution with sodium acetate, followed by repeated extraction with EtOH [McGrew & Schneider *J Am Chem Soc* 72 2547 1950]. It is a green powder that is slightly soluble in  $\text{Me}_2\text{CO}$ , EtOH and pyridine. It is soluble in conc  $\text{H}_2\text{SO}_4$  to give a blue solution that becomes turquoise in colour on dilution. [Allen et al. *J Org Chem* 7 63 1942, Beilstein 14 H 725.]

**9,10-Anthraquinone-2,6-disulfonic acid (disodium salt)** [853-68-9] M 412.3, m >325°,  $\text{pK}_{\text{Est}} \sim <0$  (for  $\text{SO}_3\text{H}$ ). Crystallise it three times from water, in the dark [Moore et al. *J Chem Soc. Faraday Trans 1* 82 745 1986]. [Beilstein 11 IV 673.]

***o*-Arsanilic acid** [2045-00-3] M 216.1, m 153°,  $\text{pK}_1^{22}$  3.77 ( $\text{AsO}_3\text{H}_2$ ),  $\text{pK}_2^{22}$  8.66 ( $\text{AsO}_3\text{H}^-$ ). Crystallise it from water or ethanol/ether. **POISONOUS.** [Beilstein 16 I 463.]

***p*-Arsanilic acid** [98-50-0] M 216.1, m 232°,  $\text{pK}_1^{22}$  4.05 ( $\text{AsO}_3\text{H}_2$ ),  $\text{pK}_2^{22}$  8.66 ( $\text{AsO}_3\text{H}^-$ ). Crystallise it from water or ethanol/ether. **POISONOUS.** [Beilstein 16 I 466.]

**Arsenazo I [3(2-arsonophenylazo)-4,5-dihydroxy-2,7-naphthalenedisulfonic acid di Na salt]** [66019-20-3] M 614.3,  $\epsilon$   $2.6 \times 10^4$  at 500nm, pH 8.0,  $\text{pK}_1$  0.6(0.8),  $\text{pK}_2$  3.52,  $\text{pK}_3$  2.97( $\text{AsO}_3\text{H}_2$ ),  $\text{pK}_4$  8.20( $\text{AsO}_3\text{H}^-$ ),  $\text{pK}_5$  9.98(OH),  $\text{pK}_6$  15.0. A saturated aqueous solution of the free acid is slowly added to an equal volume of conc HCl. The orange precipitate is filtered off, washed with acetonitrile and dried for 1-2 hours at 110° [Fritz & Bradford *Anal Chem* 30 1021 1958]. It is then titrated with NaOH to form the di or tri Na salt as set out below.

**Arsenazo III [3,6-bis(2-arsonophenylazo)-4,5-dihydroxy-2,7-naphthalenedisulfonic acid di Na salt]** [62337-00-2] M 776.4,  $\text{pK}_1$  -2.7,  $\text{pK}_2$  -2.7,  $\text{pK}_3$  0.6,  $\text{pK}_4$  0.8,  $\text{pK}_5$  1.6,  $\text{pK}_6$  3.4,  $\text{pK}_7$  6.27,  $\text{pK}_8$  9.05,  $\text{pK}_9$  11.98,  $\text{pK}_{10}$  15.1. Contaminants include monoazo derivatives, starting materials for synthesis and by-products. It is partially purified by precipitation of the dye from aqueous alkali on addition of HCl. More thorough purification is achieved by taking a 2g sample in 15-25ml of 5% aqueous  $\text{NH}_3$  and filter. Add 10ml HCl (1:1) to the filtrate to precipitate the dye. Repeat the procedure and dissolve the solid dye (0.5g) in 7ml of a 1:1:1 mixture of *n*-propanol/conc  $\text{NH}_3$ /water at 50°. After cooling, filter the solution and chromatograph the filtrate through a cellulose column using a 3:1:1 mixture of *n*-propanol/conc  $\text{NH}_3$ /water as eluent. Collect the blue band and evaporate it to 10-15ml below 80°, then add 10ml of conc HCl to precipitate pure Arsenazo III. Wash it with EtOH and dry it in air [Borak et al. *Talanta* 17 215 1970]. The sodium salt is then obtained by dissolving it in the equivalent amount of dilute NaOH and freeze-drying. The purity of the dye can be checked by paper chromatography using M HCl as eluent. It is used for the estimation of Th, U, Zr, Cd and Zn [Michaylova & Yuroukova *Anal Chim Acta* 68 73 1974.]

**Benzaldehyde-2-sulfonic acid sodium salt** [1008-72-6] M 208.2, m decomposes on heating. It forms prisms or plates by extracting with boiling EtOH, filtering, evaporating to dryness and recrystallising the Na salt from a small volume of  $\text{H}_2\text{O}$ . The *N*-phenylhydrazone sodium salt recrystallises from  $\text{H}_2\text{O}$  with m 174.5°. [Gnehm & Schüle *Justus Liebigs Ann Chem* 299 363 1898, Beilstein 11 IV 652.]

**Benzeneselenenyl bromide (phenylselenenyl bromide)** [34837-55-3] **M 236.0, m 58-62°, 60°, 62°, b 107-108°/15mm, 134°/35mm.** Distil it in a vacuum, recrystallise it from petroleum ether, CHCl<sub>3</sub> (EtOH free), or Et<sub>2</sub>O (cooling mixture) to give dark red or orange crystals. These sublime at 25°/0.001mm [Behaghel et al. *Chem Ber* **65** 815 1932, Pitteloud & Petrzilka *Helv Chim Acta* **62** 1319 1979]. [Beilstein **6** III 1111.] **HIGHLY TOXIC.**

**Benzeneselenenyl chloride (benzeneselenyl chloride, phenylselenenyl chloride)** [5707-04-0] **M 191.5, m 59-60°, 64-65°, b 92°/5mm, 120°/20mm.** Purify it by distillation in a vacuum, and recrystallisation (orange needles) from hexane [Foster *J Am Chem Soc* **55** 822 1933, Foster et al. *Rec Trav Chim Pays Bas* **53** 405, 408 1934, Behaghel & Seibert *Chem Ber* **66** 714 1933]. [Beilstein **6** III 1110.] **HIGHLY TOXIC.**

**Benzeneseleninic acid** [6996-92-5] **M 189.1, m 122-124°, pK<sup>25</sup> 4.70.** Add 10% excess of 15M NH<sub>3</sub> to the solid acid and stir until the solid dissolves, filter, decolorise with charcoal (2x, Norite) and acidify by slow addition of 6M HCl, filter the solid off and wash it with H<sub>2</sub>O. Dissolve the acid in the minimum volume of MeOH, and this solution is added dropwise to boiling H<sub>2</sub>O until cloudiness appears. At this point add 25% more boiling H<sub>2</sub>O, filter hot (decolorise if necessary) and cool rapidly, with scratching, to 0°. After 30 minutes the solid is filtered off and recrystallised as before but with very slow cooling. The colourless needles are filtered off and dried in a vacuum desiccator (CaCl<sub>2</sub>) before the melting point is measured [McCullough & Gould *J Am Chem Soc* **71** 674 1949]. The HNO<sub>3</sub> complex has **m 112°.** [Beilstein **11** H 422, **11** I 110, **11** III 716.]

**Benzeneseleninic anhydride** [17697-12-0] **M 360.1, m 124-126°, 164-165°, 170-173°.** When the anhydride is recrystallised from \*C<sub>6</sub>H<sub>6</sub> it has **m 124-126°**, but when this is heated at 140°/1 hour in a vacuum or at 90°/2 hours it has **m 164-165°** and gives a solid **m 124-126°** when then recrystallised from \*C<sub>6</sub>H<sub>6</sub>. Both depress the melting point of the acid PhSeO<sub>2</sub>H. If the high melting anhydride is dissolved in \*C<sub>6</sub>H<sub>6</sub> and seeded with the high melting anhydride, the high melting anhydride crystallises out. It readily absorbs H<sub>2</sub>O to form the acid (PhSeO<sub>2</sub>H, **m 122-124°**). Because of this, the commercial anhydride could contain up to 30% of the acid. It is best purified by converting to the HNO<sub>3</sub> complex (**m 112°**) and heating this *in vacuo* at 120°/72 hours to give the anhydride as a white powder **m 164-165°.** *Alternatively,* heat the anhydride *in vacuo* at 120°/72 hours until the IR shows no OH band. [Ayvrey et al. *J Chem Soc* 2089 1962, Barton et al. *J Chem Soc, Perkin Trans 1* 567 1977, Beilstein **11** H 422, **11** I 110, **11** III 716, **11** IV 708.] **TOXIC** solid.

**Benzeneselenol (phenylselenol, selenophenol)** [645-96-5] **M 157.1, b 57-59°/8mm, 71-72°/18mm, 84-86°/25mm, d<sub>4</sub><sup>20</sup> 1.480, n<sub>D</sub><sup>20</sup> 1.616.** Dissolve it in aqueous N NaOH, acidify this with conc HCl and extract with Et<sub>2</sub>O, dry over CaCl<sub>2</sub>, filter, evaporate on a steam bath and distil the residue from a Claisen flask or through a short column collecting the middle fraction, and seal immediately in a glass vial, otherwise the colourless liquid becomes yellow. The alkali insoluble materials consist of diphenylselenide (**b 167°/16mm**) and diphenyldiselenide, **m 63°** (from EtOH). **TOXIC**, use rubber gloves. It has a foul odour. [Foster *Org Synth Coll Vol III* 771 1955, Beilstein **6** III 1104, 1110, **6** IV 1777.]

**Benzeneselenonic acid (benzeneselenoic acid)** [39254-48-3] **M 205.1, m 64°, pK<sup>25</sup> 4.79.** Purify it by dissolving in H<sub>2</sub>O and passing through a strong cation exchange resin (H<sup>+</sup> form). Evaporate the effluent under reduced pressure and dry the residue in a high vacuum to give colourless hygroscopic crystals [Dostal et al. *Z Chem* **6** 153 1966 and IR: Dostál et al. *Chem Ber* **104** 2044 1971]. [Beilstein **11** H 422, **11** I 111.]

**Benzenestibonic acid** [535-46-6] **M 248.9, m >285°(dec).** It crystallises from acetic acid (needles), or from EtOH/CHCl<sub>3</sub> mixture on addition of water. [Schmidt *Justus Liebigs Ann Chem* **542** 288 1939, May *J Chem Soc* **101** 1033 1912.]

**Benzopurpurin 4B** {3,3'-(3,3'-dimethyl[1,1'-biphenyl]-4,4'-diyl)bis(azo)]bis[4-amino-1-naphthalene-

sulfonic acid] di-Na salt, Direct red 2} [992-59-6] M 724.7,  $\lambda_{\max}$  500nm, CI 23500,  $\text{pK}^{25} < 0$ . It crystallises from  $\text{H}_2\text{O}$ . It is a biological stain that is violet at pH 1.2 and red at pH 4.0 and is used for detecting Al, Mg, Hg, Au and U. [Beilstein 16 H 411, 16 II 224, 16 III 474, 16 IV 601.]

**Benzotriazol-1-yloxytris(dimethylamino)phosphonium hexafluorophosphate (BOP, Castro's Reagent)** [56602-33-6] M 442.29,  $m > 130^\circ(\text{dec})$ , 147-149 $^\circ(\text{dec})$ . Dissolve it in  $\text{CH}_2\text{Cl}_2$ , dry ( $\text{MgSO}_4$ ), filter, concentrate it under a vacuum, then add dry  $\text{Et}_2\text{O}$  and filter off the first crop. Add  $\text{CH}_2\text{Cl}_2$  to the filtrate and concentrate again to obtain a second crop. The solid is washed with dry  $\text{Et}_2\text{O}$  and dried in a vacuum. Also recrystallise it from dry  $\text{Me}_2\text{CO}/\text{Et}_2\text{O}$  and check the purity by NMR. Store it in the dark. [Castro et al. *Synthesis* 751 1976, Nguyen et al *J Chem Soc, Perkin Trans I* 1915 1987, Coste et al. *Tetrahedron Lett* 36 4253 1995.]

**Benzyl Orange [4-(4-benzylaminophenylazo)benzenesulfonic acid potassium salt]** [589-02-6] M 405.5,  $\text{pK}_{\text{Est}(1)} \sim < 0$ ,  $\text{pK}_{\text{Est}(2)} \sim 3.8$ . Crystallise it from  $\text{H}_2\text{O}$ .

**Benzyltriphenylphosphonium chloride** [1100-88-5] M 388.9,  $m 280^\circ$  (sintering), 287-288 $^\circ$ . Wash it with  $\text{Et}_2\text{O}$  and crystallise it from  $\text{EtOH}$  (six-sided plates). It is *hygroscopic* and forms crystals with one molecule of  $\text{H}_2\text{O}$ . [Michaelis & Soden *Justus Liebigs Ann Chem* 229 320 1885, Kröhnke *Chem Ber* 83 291 1950, Beilstein 16 IV 994.]

**Biphenyl-2-yl diphenyl phosphate** [132-29-6] M 302.4,  $b 280-290^\circ/9\text{mm}$ ,  $d_{25}^{25} 1.184$ ,  $n_D^{25} 1.5925$ . Distil the ester in a vacuum, then percolate it through an alumina column. Pass the ester through a packed column maintained at  $150^\circ$  to remove residual traces of volatile materials by a counter-current stream of nitrogen at reduced pressure. [Dobry & Keller *J Phys Chem* 61 1448 1957, Beilstein 6 III 3295, 6 IV 4585.]

**Bis(*p*-tert-butylphenyl)phenyl phosphate** [115-87-7] M 438.5,  $b 281^\circ/5.5\text{mm}$ ,  $d_4^{25} 1.108$ ,  $n_D^{25} 1.5412$ . Purify as for biphenyl-2-yl diphenyl phosphate (above). [Beilstein 6 III 1871, 6 IV 3310.]

**Bis(2-chlorophenyl) phenyl phosphate** [597-80-8] M 395,  $b 254^\circ/4\text{mm}$ ,  $n_D^{25} 1.5767$ . Purify as for biphenyl-2-yl diphenyl phosphate above. [Fox et al. *J Phys Chem* 59 1097 1955, Beilstein 6 III 680, 6 IV 807.]

**Bis(2-ethylhexyl) 2-ethylhexyl phosphonate** [25103-23-5] M 434.6,  $n_D^{25} 1.4473$ . Purify it by stirring a 0.4M solution in \*benzene with an equal volume of 6M  $\text{HCl}$  at *ca*  $60^\circ$  for 8 hours. The \*benzene layer is then shaken successively with equal volumes of water (twice), aqueous 5%  $\text{Na}_2\text{CO}_3$  (three times), and water (eight times), followed by evaporation of the \*benzene and distillation of the residue under reduced pressure at room temperature (using a rotating evacuated flask). It should be stored dry and in the dark [Peppard et al. *J Inorg Nucl Chem* 24 1387 1962]. Distil it in a vacuum, then percolate it through an alumina column before finally passing it through a packed column maintained at  $150^\circ$  where residual traces of volatile materials are removed by a counter-current stream of  $\text{N}_2$  under reduced pressure [Dobry & Keller *J Phys Chem* 61 1448 1957].

**Bis(2-ethylhexyl) phosphoric acid ('diisooctyl' phosphate)** [298-07-7, 27215-10-7] M 322.4,  $m -60^\circ$ ,  $b 209^\circ/10\text{mm}$ ,  $d_4^{20} 0.965$ ,  $\text{pK}_{\text{Est}} \sim 1.7$ . Contaminants of commercial samples include the monoester, polyphosphates, pyrophosphate, 2-ethylhexanol and metal impurities. Dissolve the acid in *n*-hexane to give an 0.8M solution. Wash this with an equal volume of M  $\text{HNO}_3$ , then with saturated  $(\text{NH}_4)_2\text{CO}_3$  solution, with 3M  $\text{HNO}_3$ , and twice with water [Petrow & Allen *Anal Chem* 33 1303 1961]. Similarly, the impure sodium salt, after scrubbing with petroleum ether, is acidified with  $\text{HCl}$  and the free organic acid is extracted into petroleum ether and purified as above [Peppard et al. *J Inorg Nucl Chem* 7 231 1958], or as described by Stewart & Crandall [*J Am Chem Soc* 73 1377 1951]. It can be purified *via* its copper salt [McDowell et al. *J Inorg Nucl Chem* 38 2127 1976]. [Beilstein 1 IV 1796.]

**2,4-Bis(methylthio)-1,3,2 $\lambda^5$ ,4 $\lambda^5$ -dithiadiphosphetane-2,4-dithione (Davy's reagent methyl)** [82737-61-9] M 284.4,  $m 160^\circ$ . It crystallises from \* $\text{C}_6\text{H}_6$  in yellow plates or from hot trichlorobenzene. The low melting point reported in the literature ( $112^\circ$  with gradual softening at  $68-102^\circ$ ) has been attributed to the presence of elemental sulfur in the crystals. It has a **foul odour** and is a *suspected carcinogen*. [Yousif et al. *Tetrahedron*

40 2663 1984, Scott et al. *J Org Chem* 22 789 1957.]

**Bismuthiol I (2,5-dimercapto-1,3,4-thiadiazole) potassium salt** [4628-94-8] **M 226.4, m 275-276°(dec), ~294°(dec), pK<sub>Est(1)</sub>~4.1.** Usually contaminated with disulfide. Purify it by recrystallisation from EtOH. It is a reagent for the detection of Bi, Cu, Pb and Sb. [Najumdar & Chacraborty *Anal Chim Acta* 19 372 1958, Fries *Z Anal Chem* 165 100, 1959, Beilstein 27 IV 7725.]

**Bis[4-(1,1,3,3-tetramethylbutyl)phenyl]phosphate calcium salt (Selectophore)** [40835-97-0] **M 987.3.** The Ca diester salt is washed with H<sub>2</sub>O (x3) and MeOH (x3) alternately and dried in a vacuum oven at 50°. If the Ca salt is contaminated with much Ca salt of the monoester, then it (10g) is converted to the free acid by adding 6N HCl (*ca* 10 volumes) and Et<sub>2</sub>O (> 50 volumes) is added to it and stirred vigorously to form the free acids. When no white precipitate remains (*ca* 5 minutes), the Et<sub>2</sub>O is separated, washed with H<sub>2</sub>O (2 >50 ml) and dried by filtering through a bed of anhydrous Na<sub>2</sub>SO<sub>4</sub> (11 x 5 cm) which is then washed with Et<sub>2</sub>O (2 >50 ml). Evaporation gives an oil (TLC R<sub>F</sub> 0.81 for diester and 0.50 for monoester). The oil is dissolved in \*benzene (*ca* 25ml) and extracted with ethane-1,2-diol (25ml, 10x). After ten washings, a small sample of the \*benzene layer is washed twice with H<sub>2</sub>O to remove the diol and showed that it is pure *bis*-[4-(1,1,3,3-tetramethylbutyl)-phenyl]phosphoric acid by TLC, i.e. no monophosphate. To form the Ca salt, the oil is dissolved in MeOH and to it is added the equivalent amount of CaCl<sub>2</sub> together with aqueous NaOH to keep the pH >10. The resulting white precipitate is collected, washed alternately with 3 batches of H<sub>2</sub>O and MeOH and dried in a vacuum oven at 50°. [Craggs et al. *J Inorg Nucl Chem* 40 1483 1978, Morton & Chung *Anal Biochem* 157 345 1986.]

**2,4-Bis(*p*-tolylthio)-1,3,2λ<sup>5</sup>,4λ<sup>5</sup>-dithiadiphosphetane-2,4-dithione (Heimgartner's reagent or Davy's reagent *p*-tolyl)** [114234-09-2] **M 436.6, m 175-176°.** Recrystallise it from toluene (light yellow solid), wash it with Et<sub>2</sub>O and dry *in vacuo*. [Jennt & Heimgartner *Helv Chim Acta* 70 1001 1987.]

***N,O*-Bis(trimethylsilyl)acetamide (BSA)** [10416-59-8] **M 203.4, b 71-73°/35mm, d<sub>4</sub><sup>20</sup> 0.836, n<sub>D</sub><sup>20</sup> 1.4150.** Fractionate it through a spinning band column and collect liquid **b** 71-73°/35mm, and not higher because the main impurity MeCONHSiMe<sub>3</sub> distils at **b** 105-107°/35mm. It is used for derivatising alcohols and sugars [Klebe et al. *J Am Chem Soc* 88 3390 1966, see Matsuo et al. *Carbohydr Res* 241 209 1993, Johnson *Carbohydr Res* 237 313 1992]. It is **FLAMMABLE** and **TOXIC**.

**Bis(trimethylsilyl)acetylene (BTMSA)** [14630-40-1] **M 170.4, m 26°, b 134-136°/atm.** Dissolve it in petroleum ether and wash it with ice-cold dilute HCl. The petroleum ether extract is dried (MgSO<sub>4</sub>), evaporated and fractionated at 760mm. [Walton & Waugh *J Organomet Chem* 37 45 1972, Beilstein 4 IV 3950.]

**Bis(trimethylsilyl) sulfide (hexamethyldisilathiane)** [3385-94-2] **M 178.5, b 65-67°/16mm, 162.5-163.5°/750mm, 164°/760mm, d<sub>4</sub><sup>20</sup> 0.85, n<sub>D</sub><sup>20</sup> 1.4598.** Dissolve it in petroleum ether (*b ca* 40°), remove the solvent and distil it. Redistil it under atmospheric pressure of dry N<sub>2</sub>. It is collected as a colourless liquid which solidifies to a white solid in Dry-ice. On standing for several days it turns yellow possibly due to liberation of sulfur. Store it below 4° under dry N<sub>2</sub>. [Eaborn *J Chem Soc* 3077 1950, Beilstein 4 IV 4033.]

**9-Borabicyclo[3.3.1]nonane (9BBN)** [*monomer* 280-64-8] [*dimer* 21205-91-4 or 70658-61-6] [*1:1 coordination compound with tetrahydrofuran* 76422-63-4] **M 122.0 (monomer), 244.0 (dimer), m 141-143° (monomer), 150-152°, 154-155° (dimer), b 195°/12mm.** It is available as the solid dimer or in tetrahydrofuran solution. The solid is relatively stable and can be purified by distillation in a vacuum (as dimer) and by recrystallisation from tetrahydrofuran (solubility at room temperature is 9.5%, 0.78M), filter off the solid under N<sub>2</sub>, wash it with dry pentane and dry it *in vacuo* at *ca* 100°. The solid is a dimer (IR 1567cm<sup>-1</sup>), stable in air (for *ca* 2 months), and can be heated for 24 hours at 200° in an inert atmosphere without loss of hydride activity. It is a dimer in tetrahydrofuran solution also (IR 1567cm<sup>-1</sup>). It is sensitive to H<sub>2</sub>O and air (O<sub>2</sub>) in solution. Its concentration in solution can be determined by reaction with MeOH and measuring the volume of H<sub>2</sub> liberated, or it can be oxidised to *cis*-cyclooctane-1,5-diol (**m** 73.5-74.5°). [IR: Knights & Brown *J Am Chem Soc* 90 5280 1968, Brown et al. *J Am Chem Soc* 96 7765 1974, Brown et al. *J Org Chem* 41 1778 1976, Brown & Chen *J Org*

*Chem* **46** 3978 1981, Fieser & Fieser's *Reagents for Org Synth* **2** 31, **3** 24, **10** 48, **15** 43, **17**, 49.]

**Borane pyridine complex** [110-51-0] **M 92.9, m 8-10°, 10-11°, b 86°/7mm, 100-101°/12mm,  $d_4^{20}$  0.785.** Dissolve it in Et<sub>2</sub>O and wash it with H<sub>2</sub>O in which it is insoluble. Evaporate the Et<sub>2</sub>O and distil the residual oil to gives better than 99.8% purity. Its vapour pressure is less than 0.1mm at room temperature. [Taylor et al. *J Am Chem Soc* **77** 1506 1955, *Beilstein* **20** IV 2235.]

**Borane triethylamine complex** [1722-26-5] **M 115.0, b 76°/4mm, 8°/7mm, 100-101°/12mm,  $d_4^{20}$  0.78.** Distil it in a vacuum using a 60cm glass helices-packed column. [Brown et al. *J Am Chem Soc* **64** 325 1942, Ashby & Foster *J Am Chem Soc* **84** 3407 1962, Matsuura & Tolcura *Tetrahedron Lett* 4703 1968, *Beilstein* **4** IV 329.]

**Borane trimethylamine complex** [75-22-9] **M 73.0, m 94-94.5°, b 171°/atm.** It is sublimed using equipment described in Burg and Schlesinger [*J Am Chem Soc* **59** 780 1937]. Its vapour pressure is 86mm at 100°. It forms colourless hexagonal crystals varying from needles to short lumps, which are slightly soluble in H<sub>2</sub>O (1.48% at 30°), EtOH (1%), hexane (0.74%), but very soluble in Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> and AcOH. It is stable at 125°. [Burg & Schlesinger *J Am Chem Soc* **59** 780 1937, Brown et al. *J Am Chem Soc* **104** 325 1942, *Beilstein* **4** IV 140.]

**2-Bromoallyltrimethylsilane** [81790-10-5] **M 193.2, b 64-66°/10mm, 82-85°/58-60mm,  $d_4^{25}$  1.13.** It is fractionally distilled through an efficient column. It is **flammable**. [Trost & Chan *J Am Chem Soc* **104** 3733 1982, Trost & Coppola *J Am Chem Soc* **104** 6879 1982.]

**2-Bromo-1,3,2-benzodioxaborole** [51901-85-0] **M 198.8, m 47°, 51-53°, b 76°/9mm.** Keep at 20°/15mm for some time and then fractionally distil. [Gerrard *J Chem Soc* 1529 1959, *Beilstein* **6** IV 5612.]

**IR(endo, anti)-3-Bromocamphor-8-sulfonic acid ammonium salt** [55870-50-3] **M 328.2, m 284-285°(dec),  $[\alpha]_D^{25} +84.8^\circ$  (c 4, H<sub>2</sub>O).** Pass a hot aqueous solution of it through an alumina column to remove water-soluble coloured impurities which remain on the column when the ammonium salt is eluted with hot water. The salt is crystallised from water and is dried over CaCl<sub>2</sub> in a desiccator [Craddock & Jones *J Am Chem Soc* **84** 1098 1962, Kauffmann *J Prakt Chem* **33** 295 1966]. [*Beilstein* **11** H 319, **11** I 77, **11**, II 183, **11** III 595.]

**Bromopyrogallol Red.** See in "Aromatic Compounds", in this Chapter.

**Bromosulfalein (phenoltetrabromophthalein 3',3'-disulfonic acid disodium salt)** [71-67-0] **M 838.0.** Purify it by TLC on silica Gel G (Merck 250 $\mu$  particle size) in two solvent systems (BuOH/AcOH/H<sub>2</sub>O 30:7.5:12.5 v/v, and BuOH/propionic acid/H<sub>2</sub>O 30:20:7.5 v/v). When the solvent reaches a height of ~10cm, the plate is removed, dried in air and developed with NH<sub>3</sub> vapour giving blue-coloured spots. Also, the dye can be chromatographed on MN Silica Gel with *t*-BuOH/H<sub>2</sub>O/*n*-BuOH (32:10:5 v/v) as eluent and visualised with a dilute KOH (or NaOH if the Na salt is required) spray. The product corresponding to bromosulfalein is scraped off and eluted with H<sub>2</sub>O, filtered and evaporated to dryness in a vacuum. It is then dissolved in H<sub>2</sub>O, filtered through Sephadex G-25 and evaporated to dryness. [UV and IR identification: Barbier & DeVeerd *J Pharm Sci* **57** 819 1968, NMR: Kato et al. *Chem Pharm Bull Jpn* **20** 581 1972, McGuire *Anal Biochem* **83** 75 1977, *Beilstein* **18/9** V 461.]

**Bromotrimethylsilane (trimethylbromosilane, trimethylsilyl bromide)** [2857-97-8] **M 153.1, m -43.5° to -43.2°, b 40.5°/200mm, 77.3°/735mm, 79°/744mm, 79.8-79.9°/754mm,  $d_4^{20}$  1.1805,  $n_D^{20}$  1.422.** Purify it by repeated fractional distillation and store it in sealed ampoules in the dark. [McCusker & Reilly *J Am Chem Soc* **75** 1583 1953.] Also fractionate it through a 15-plate column (0.8 x 32cm packed with 1/16in single turn helices of Pt-Ir wire). [Gilliam et al. *J Am Chem Soc* **68** 1161 1946, Pray et al. *J Am Chem Soc* **70** 433 1948, *Beilstein* **4** IV 4008.]

**But-3-enylboronic acid** [379669-72-4] **M 99.9, m 84-90°, pK<sub>Est</sub> 8.8.** Recrystallise the acid from toluene and dry it *in vacuo*. [*cf* Letsinger & Skoog *J Org Chem* **18** 895 1953.]

**Butylboronic acid (1-butanedihydroxyborane)** [4426-47-5] M 101.9, m 90-92°, 94-96°, pK<sub>Est</sub> ~8.8. Purify the acid by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and dry it *in vacuo*. [Corey et al. *J Am Chem Soc* **116** 3151 1994, Quallich et al. *J Am Chem Soc* **116** 8515 1994, Seerden *Tetrahedron Lett* **35** 4419 1994, *Beilstein* **4** IV 4383.]

**(±)-sec-Butylboronic acid ([sec-butyl]-dihydroxyborane)** [88496-88-2] M 101.9, m 86-89°, 87-88°, pK<sub>Est</sub> ~8.8. Purify the acid by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and dry *in vacuo*. [McCusker et al. *J Am Chem Soc* **79** 5179 1957, *Beilstein* **4** IV 4386.]

**tert-Butyldicyclohexylphosphine (dicyclohexyl-tert-butylphosphine, Cy<sub>2</sub>P<sup>t</sup>Bu)** [93634-87-8] M 254.4, m ~22-25°, d<sub>4</sub><sup>25</sup> 1.094, pK<sub>Est</sub> ~8.7. This phosphine was prepared by adding *t*-BuLi (18ml, 27.2mmol, 1.5M in pentane) dropwise to a solution of chloro-dicyclohexylphosphine (5.75g, 24.7mmol, [16523-54-9] see above) in THF (20ml) at -78° and the yellow suspension was allowed to warm to ~25° and stirred overnight. The mixture was evaporated to dryness, extracted with pentane (2 x 20ml), filtered through Celite, concentrated to 20ml, cooled to -78° and after 2 hours the Cy<sub>2</sub>P<sup>t</sup>Bu crystallised in white crystals (5.6g, 89%) which were filtered off and washed with cold pentane under N<sub>2</sub> or Ar. It is highly air sensitive, flammable and should be stored under N<sub>2</sub> or Ar. It melts at room temperature. The <sup>1</sup>H NMR (CDCl<sub>3</sub>) has peaks at δ: 1.90-1.05 (m, 22H, C<sub>6</sub>H<sub>11</sub>), 1.10 (d, 9H, C(CH<sub>3</sub>)<sub>3</sub>, <sup>3</sup>J<sub>H-P</sub> = 10.8Hz); the <sup>13</sup>C NMR (CDCl<sub>3</sub>) has peaks at δ: 33.11 (d, CMe<sub>3</sub>, <sup>1</sup>J<sub>C-P</sub> = 19.4Hz), 33.61 (d, C<sub>1</sub> C<sub>6</sub>H<sub>11</sub>, <sup>1</sup>J<sub>C-P</sub> = 16.2Hz), 30.96, 27.82 (2d, C<sub>2</sub> C<sub>6</sub>H<sub>11</sub>, <sup>2</sup>J<sub>C-P</sub> = 9.7Hz), 30.33 (d, CH<sub>3</sub>)<sub>3</sub>, <sup>2</sup>J<sub>C-P</sub> = 13.2Hz), 27.69, 27.61 (2s, C<sub>3</sub> C<sub>6</sub>H<sub>11</sub>), 26.41 (s, C<sub>4</sub> C<sub>6</sub>H<sub>11</sub>); and <sup>31</sup>P NMR (CDCl<sub>3</sub>) has a peak at δ 28.58. [Jan et al. *J Organomet Chem* **55** 606 2000.]

**tert-Butyldimethylsilyl chloride (TBDMSCl)** [18162-48-6] M 150.7, m 87-89°, 92.5°, b 125°/760mm. Fractionally distil it at atmospheric pressure. [Sommer & Tyler *J Am Chem Soc* **76** 1030 1954, Corey & Venkateswarlu *J Am Chem Soc* **94** 6190 1972, *Beilstein* **4** IV 4076.]

**tert-Butyldiphenylchlorosilane (TBDPSCl, tert-butylchlorodiphenylsilane)** [58479-61-1] M 274.9, b 90°/0.015mm, d<sub>4</sub><sup>20</sup> 1.057, n<sub>D</sub><sup>20</sup> 1.568. Purify it by repeated fractional distillation. It is soluble in DMF and pentane [Hanessian & Lavalee *Can J Chem* **53** 2975 1975, Robl et al. *J Med Chem* **34** 2804 1991]. [*Beilstein* **4** IV 4076 for *tert-butylchlorodimethylsilane*.]

***n*-Butylphenyl *n*-butylphosphonate** [36411-99-1] M 270.3. Crystallise it three times from hexane as its compound with uranyl nitrate. See *tri-n-butyl phosphate* below.

***p*-tert-Butylphenyl diphenyl phosphate** [981-40-8] M 382.4, b 261°/6mm, n<sub>D</sub><sup>25</sup> 1.5522. Purify it by vacuum distillation, and percolation through an alumina column, followed by passage through a packed column maintained at 150° to remove residual traces of volatile materials in a counter-current stream of N<sub>2</sub> at reduced pressure [Dobry & Keller *J Phys Chem* **61** 1448 1957].

**Cacodylic acid (dimethylarsinic acid)** [75-60-5] M 138.0, m 195-196°, pK<sub>1</sub><sup>25</sup> 1.57, pK<sub>2</sub><sup>25</sup> 6.27 [Me<sub>2</sub>As(O)OH]. Recrystallise it from warm EtOH (3ml/g) by cooling and filtering. Dry it in a vacuum desiccator over CaCl<sub>2</sub>. It has also been recrystallised twice from propan-2-ol. [Kilpatrick *J Am Chem Soc* **71** 2608 1949, Nichol *J Am Chem Soc* **72** 2367 1950, Koller & Hawkrigde *J Am Chem Soc* **107** 7412 1985, *Beilstein* **4** H 610, **4** I 567, **4** II 993, **4** III 1818, **4** IV 3681.]

**Cadion [1-(4-nitrophenyl)-3-(4-phenylazophenyl)-triazene]** [5392-67-6] M 346, m 189°(dec). Commercial cadion is purified by recrystallisation from 95% EtOH and is dried *in vacuo*. It is stable in 0.2 N KOH (in 20% aqueous EtOH) at 25°. It is a sensitive reagent for Cd, and the Cd complex has λ<sub>max</sub> (EtOH) 475nm. [Chavanne & Geronimi *Anal Chim Acta* **19** 377 1958, *Beilstein* **16** III 664.]

**(4-Carbamylphenylarsylenedithio)diacetic acid** [531-72-6] M 345.1, pK<sub>Est</sub> ~3.5. Recrystallise it from MeOH or EtOH.

**Catecholborane (1,3,2-Benzodioxaborole)** [274-07-7] M 119.2, b 50°/50mm, 66°/80mm, 76-77°/100mm, 88°/165mm,  $d_4^{20}$  1.125,  $n_D^{20}$  1.507 (also available as a 1.0M solution in THF or toluene). It is a moisture-sensitive flammable liquid which is purified by distillation in a vacuum under a N<sub>2</sub> atmosphere and stored under N<sub>2</sub> at 0-4°. It liberates H<sub>2</sub> when added to H<sub>2</sub>O or MeOH. A solution in THF, after 25 hours at 25°, has residual hydride of 95% (under N<sub>2</sub>) and 80% (under air) [Brown & Gupta *J Am Chem Soc* 97 5249 1975].

**Chloramine-T (N-chloro-p-toluenesulfonamide sodium salt) 3H<sub>2</sub>O** [7080-50-4] M 281.7, m 168-170°(dec). Recrystallise it from hot water (2ml/g). Dry it in a desiccator over CaCl<sub>2</sub> where it loses water. Protect it from sunlight. It is used for the detection of bromate and halogens, and Co, Cr, Fe, Hg, Mn, Ni and Sb ions. [Campbell & Johnson *Chem Rev* 78 65 1978, Bremner *Synthetic Reagents* 6 9 1985, Chattaway *J Chem Soc* 87 145 11905, Inglis *J Soc Chem Ind (Lond)* 37 288 1918, *Beilstein* 11 H 107, 11 I 29, 11 II 62, 11 III 300, 2 IV 457.]

**Chlorazol Sky Blue FF {6,6'-[(3,3'-dimethoxy[1,1'-biphenyl]-4,4'-diyl)bis(azo)bis(4-amino-5-hydroxy-1,3-naphthylenedisulfonic acid) tetra-Na salt** [2610-05-1] M 996.9, m >300°(dec). Free it from other electrolytes by adding aqueous sodium acetate to a boiling solution of the dye in distilled water. After standing, the salted-out dye is filtered on a Büchner funnel, the process being repeated several times. Finally, the precipitated dye is boiled several times with absolute EtOH to wash out any sodium acetate, then dried (as the sodium salt) at 105°. [McGregor et al. *Trans Faraday Soc* 58 1045 1962, *Beilstein* 16 II 259.]

**Chlorodicyclohexylphosphine [Cy<sub>2</sub>PCI]** [16523-54-9] M 232.7, b 132-138°/3mm, 165°/12mm, 173-174°/17mm, 182-183°/23mm,  $d_4^{25}$  1.054,  $n_D^{20}$  1.533. Cy<sub>2</sub>PCI can be obtained as a colourless oil from cyclohexyldichloro-phosphine and cyclohexylmagnesium chloride, also from dicyclohexyldiethylamino-phosphine and HCl in the presence of NH<sub>4</sub>Cl in petroleum ether (b 70-90°) followed by fractional distillation. Alternatively, reaction of cyclohexylmagnesium chloride (from 12.6g of Mg and 62g of cyclohexyl chloride) in Et<sub>2</sub>O (250ml) and PCl<sub>3</sub> (35g) in Et<sub>2</sub>O (300ml) under N<sub>2</sub> followed by fractional distillation gives (22g, 37.2%) of Cy<sub>2</sub>PCI. [Issleib & Seidel *Chem Ber* 92 2681 1959, *Beilstein* 16 IV 968.]

**Chlorodiphenylphosphine (diphenylphosphinous chloride)** [1079-66-9] M 220.6, m 15-16°, b 124-126°/0.6mm, 174°/5mm, 320°/atm,  $d_4^{20}$  1.229,  $n_D^{20}$  1.636. This air-sensitive, pale yellow lachrymatory liquid is purified by careful fractional distillation and discarding the lower boiling fraction which contains the main impurity PhPCl<sub>2</sub> (b 48-51°/0.7mm), and checking for impurities by NMR. [Weinberg *J Org Chem* 40 3586 1975, Honer et al. *Chem Ber* 94 2122 1961, *Beilstein* 16 IV 969.]

**Chlorodi(o-tolyl)phosphine** [36042-94-1] M 248.7, m 36-37°, b 120-122°/0.03mm, 146-147°/1.1mm. It is purified by fractional distillation in a vacuum (b 179-183°/7mm, 253-257°/15mm) and the distillate solidifies (m 36°, also reported is m 37°). [Weinberg *J Org Chem* 40 3586 1975, McEwen et al. *J Am Chem Soc* 100 7304 1978, *Beilstein* 16 H 769, 16 IV 970 for chlorodi(p-tolyl)phosphine.]

**(Chloromethyl)dimethylvinylsilane** [16709-86-7] M 134.7, b 122-126°/atm,  $d_4^{25}$  0.908,  $n_D^{20}$  1.440. Distil the silane in a vacuum, but if it is suspect then dissolve it in Et<sub>2</sub>O, shake it with saturated aqueous NH<sub>4</sub>Cl, dry the Et<sub>2</sub>O layer (anhydrous Na<sub>2</sub>SO<sub>4</sub>), filter evaporate and fractionate in a vacuum. [Altamura et al. *J Org Chem* 60 8403 1995.]

**Chloromethylphosphonic acid dichloride** [1983-26-2] M 167.4, b 50°/0.5mm, 52-53(59)°/2mm, 63-65°/3mm, 78-79°/10mm, 87-88°/15mm, 102-103°/30mm,  $d_4^{20}$  1.638,  $n_D^{20}$  1.4971. It is fractionally distilled using a short Claisen column and redistilled. The aniline salt has m 199-201°. The <sup>31</sup>P NMR has a single peak at -38±2 ppm from 85% H<sub>3</sub>PO<sub>4</sub>. [Kinnear & Perren *J Chem Soc* 3437 1952, NMR: van Wazer et al. *J Am Chem Soc* 78 5715 1956, McConnell et al. *J Org Chem* 22 462 1957, *Beilstein* 1 III 2593, 1 IV 3068.]

**2-Chloro-2-oxo-1,3,2-dioxaphospholane** [6609-64-9] M 142.5, m 12-14°, b 89-91°/0.8mm,  $d_4^{20}$  1.549,  $n_D^{20}$

**1.448.** It should be distilled under high vacuum as some polymerisation occurs at atmospheric pressure. It has IR bands at 3012, 2933, 1477, 1366, 1325, 1040, 924 and 858  $\text{cm}^{-1}$ . It is hydrolysed to  $\text{HOCH}_2\text{CH}_2\text{OPO}_3\text{H}_2$  in 30 minutes in  $\text{H}_2\text{O}$  at  $100^\circ$  [IR: Cox & Westheimer *J Am Chem Soc* **80** 5441 1958]. [Beilstein **1** IV 2419.]

**2-Chlorophenyl diphenyl phosphate** [115-85-5] **M 360.7, b 236<sup>o</sup>/4mm, n<sub>D</sub><sup>25</sup> 1.5707.** Purify it by vacuum distillation, percolate it through a column of alumina, then pass it through a packed column maintained by a countercurrent stream of  $\text{N}_2$  at reduced pressure [Dobry & Keller *J Phys Chem* **61** 1448 1957].

**Chlorosulfonic (chlorosulfuric) acid** [7790-94-5] **M 116.5, b 151-152<sup>o</sup>/750mm, d<sub>4</sub><sup>20</sup> 1.753, n<sub>D</sub><sup>25</sup> 1.4929, pK<sup>25</sup> -5.9 (aqueous  $\text{H}_2\text{SO}_4$ ).** Distil it in an all-glass apparatus, taking the fraction boiling at 156-158<sup>o</sup>. It reacts **EXPLOSIVELY** with water [Cremlyn *Chlorosulfonic acid: A Versatile Reagent*, Royal Society of Chemistry UK, 2002, 308 pp, ISBN 0854044981]. **LACHRYMATORY and CORROSIVE, wear gloves and face shield.**

**Chlorotriphenylsilane (triphenylchlorosilane)** [76-86-8] **M 294.9, m 90-92<sup>o</sup>, 91-93<sup>o</sup>, 94-95<sup>o</sup>, 97-99<sup>o</sup>, b 156<sup>o</sup>/1mm, 161<sup>o</sup>/0.6mm.** Likely impurities are tetraphenylsilane, small amounts of hexaphenyldisiloxane and traces of triphenylsilanol. Purify it by distillation at 2mm, then crystallise it from EtOH-free  $\text{CHCl}_3$ , and from petroleum ether (b 30-60<sup>o</sup>) or hexane by cooling in a Dry-ice/acetone bath. [Allen & Modena *J Chem Soc* 3671 1957, Curran et al. *J Am Chem Soc* **72** 4471 1950, Speier & Zimmerman *J Am Chem Soc* **77** 6395 1955, Thomas & Rochow *J Am Chem Soc* **79** 1843 1957, Beilstein **16** IV 1484.]

**Chromazurol S (Mordant Blue 29)** [1667-99-8] **M 539.3,  $\lambda_{\text{max}}$  540nm,  $\epsilon$  7.80 x 10<sup>4</sup> (10M HCl), CI 43825, pK<sub>1</sub><sup>25</sup> <0, pK<sub>2</sub><sup>25</sup> 2.25, pK<sub>3</sub><sup>25</sup> 4.88, pK<sub>4</sub><sup>25</sup> 11.75.** The crude *phenolic triphenylmethanecarboxysulfonic acid triNa salt* (40g) is dissolved in water (250ml) and filtered. Then conc HCl (50ml) is added to the filtrate, with stirring. The precipitate is filtered off, washed with HCl (2M) and dried. It is redissolved in water (250ml), and precipitation is repeated twice more in a water bath at 70<sup>o</sup>. It is dried under vacuum over solid KOH (first), then  $\text{P}_2\text{O}_5$  [Martynov et al. *Zh Analyt Khim* **32** 519 1977]. It has also been purified by paper chromatography using *n*-butanol, acetic acid and water (7:3:1). First and second spots were extracted. It chelates Al and Be. It is used also for estimating fluoride. [Beilstein **11** IV 707.]

**Copper (I) thiophenolate** [1192-40-1] **M 172.7, m ca 280<sup>o</sup>, pK<sub>1</sub><sup>25</sup> 6.62 (for  $\text{PhS}^-$ ).** The Cu salt can be extracted from a thimble (Soxhlet) with boiling MeOH. It is a green-brown powder that gives a yellow-green solution in pyridine. Wash it with EtOH and dry it in a vacuum. It can be precipitated from a pyridine solution by adding  $\text{H}_2\text{O}$ , collecting the precipitate, washing it with EtOH and drying in a vacuum. [Posner et al. *Synthesis* 662 1974, Krebs et al. *Chem Ber* **90** 425 1957, Beilstein **6** IV 1465.]

**Cupferron ammonium salt (N-nitroso-N-phenylhydroxylamine ammonium salt)** [135-20-6] **M 155.2, m 150-155<sup>o</sup>(dec), 162.5-163.5<sup>o</sup>, 163-164<sup>o</sup>, pK<sup>25</sup> 4.16 (free base).** Recrystallise it twice from EtOH after treatment with Norite and finally once with EtOH. The crystals are washed with diethyl ether and air dried, then stored in the dark over solid ammonium carbonate. A standard solution (ca 0.05M prepared in air-free  $\text{H}_2\text{O}$ ) is prepared daily from this material for analytical work and is essentially 100% pure. [Olsen & Elving *Anal Chem* **26** 1747 1954.] It can also be washed with  $\text{Et}_2\text{O}$ , dried and stored as stated. In a sealed, dark container it can be stored for at least 12 months without deterioration. The UV has  $\lambda_{\text{max}}$  at 260nm ( $\text{CHCl}_3$ ). [Marvel *Org Synth Coll Vol I* 177 1941, Elving & Olson *J Am Chem Soc* **78** 4206 1956, Beilstein **16** IV 891.] Possible **CARCINOGEN**.

**Cupric trifluoromethylsulfonate (copper II triflate)** [34946-82-2] **M 361.7, pK<sup>25</sup> <-3.0 (for triflic acid).** Dissolve it in MeCN, add dry  $\text{Et}_2\text{O}$  until cloudy and cool at -20<sup>o</sup> in a freezer. The light blue precipitate is collected and dried in a vacuum oven at 130<sup>o</sup>/20mm for 8 hours. It has  $\lambda_{\text{max}}$  at 737nm ( $\epsilon$  22.4  $\text{M}^1\text{cm}^{-1}$ ) in AcOH. [Salomon & Kochi *J Am Chem Soc* **95** 330 1973]. It has also been dried in a vessel at 0.1Torr by heating with a Fischer burner [Andrist et al. *J Org Chem* **43** 3422 1978]. It has been dried at 110-120<sup>o</sup>/5mm for 1 hour before use and forms a \*benzene complex which should be handled in a dry box because it is air sensitive [Kobayashi et al. *Chem Pharm Bull Jpn* **28** 262 1980, Salomon & Kochi *J Am Chem Soc* **95** 330 1973]. [Beilstein **3** IV 34.]

**Cuprous (I) bromide dimethylsulfide complex** [54678-23-8] **M 205.6, m ca 135°(dec)**. Purify it by recrystallisation in the presence of Me<sub>2</sub>S. A solution of the complex (1.02g) in Me<sub>2</sub>S (5ml) is slowly diluted with hexane (20ml), and the pure colourless prisms of the complex (0.96g) separate and are collected and dried, **m 124-129°(dec)**. The complex is insoluble in hexane, Et<sub>2</sub>O, Me<sub>2</sub>CO, CHCl<sub>3</sub> and CCl<sub>4</sub>. It dissolves in DMF and DMSO, but the solution becomes hot and green indicating decomposition. It dissolves in \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O, MeOH and CHCl<sub>3</sub> if excess of Me<sub>2</sub>S is added and a colourless solution is obtained. [House et al. *J Org Chem* **40** 1460 1975.] Prior to use, the complex is dissolved in Me<sub>2</sub>S and evaporated to dryness in the weighed reaction flask [Bourgain-Commerçon et al. *J Organomet Chem* **228** 321 1983]. If this reagent is to assist in methylations with MeMgBr (e.g. of fullerenes) it is always best to make a fresh preparation. CuBr (25g) is washed with MeOH (4 x 50ml) to remove coloured impurities and dried *in vacuo* for 1 hour. This white green-tinged CuBr is dissolved in Me<sub>2</sub>S (60ml, redistilled b 38°, use efficient fume cupboard) and insoluble impurities are filtered off. Hexane (200ml) is added and the precipitated light-sensitive white crystalline complex is filtered off and washed with hexane (5 times) under suction and N<sub>2</sub>. It should be dried under a stream of dry N<sub>2</sub>, and should not be kept or stored under reduced pressure as it will lose the some of the disulfide ligand and its efficiency. [Wuts *Synth Commun* **11** 139 1981, Nakamura et al. *Org Synth* **83** 80, 2006.]

**Cuprous iodide trimethylphosphite** [34836-53-8] **M 314.5, m 175-177°, 192-193°**. Cuprous iodide dissolves in a \*C<sub>6</sub>H<sub>6</sub> solution containing trimethylphosphite to form the complex. The complex crystallises from \*C<sub>6</sub>H<sub>6</sub> or petroleum ether. [Arbusoff *Chem Ber* **38** 1171 1905, Nishizawa *Bull Chem Soc Jpn* **34** 1170, 1177 1961.]

**Decaborane (14) (B<sub>10</sub>H<sub>14</sub>)** [17702-41-9] **M 122.2, m 99.6-99.7°, 99.7-100°, b 100°/19mm, 213°/atm, d<sub>4</sub><sup>25</sup> 0.94**. Purify decaborane by vacuum sublimation at 80°/0.1mm, followed by crystallisation from methylcyclohexane, CH<sub>2</sub>Cl<sub>2</sub>, or dry olefin-free-*n*-pentane, the solvent being subsequently removed by storing the crystals in a vacuum desiccator containing CaCl<sub>2</sub>. It is soluble in H<sub>2</sub>O but is slowly decomposed to give H<sub>2</sub>. It is soluble in alkali, and on acidification it liberates H<sub>2</sub>. **TOXIC**. [Greenwood in *Comprehensive Chemistry (Ed Bailer et al.)* Pergamon Press Vol **1** pp 818-837 1973.]

**Di-*n*-amyl *n*-amylphosphonate** [6418-56-0] **M 292.4, b 150-151°/2mm, n<sub>D</sub><sup>20</sup> 1.4378**. Purify it by three crystallisations of its uranyl nitrate complex from hexane (see *tributyl phosphate*). It extracts Zr<sup>2+</sup> from NaCl solutions.

**Di-*n*-butyl boron triflate (di-*n*-butylboryl trifluoromethanesulfonate)** [60669-69-4] **M 274.1, b 37°/0.12mm, 60°/2mm, pK<sup>25</sup> <-3.0 (for triflic acid)**. Distil it in a vacuum under argon and store it under argon. It should be used within 2 weeks of purchase or after redistillation. Use a short path distillation system. It has IR bands in CCl<sub>4</sub> with  $\nu_{\max}$  at 1405, 1380, 1320, 1200 and 1550cm<sup>-1</sup>, and <sup>13</sup>C NMR (CDCl<sub>3</sub>) with  $\delta$  at 118.1, 25.1, 21.5 and 13.6. [Gage & Evans *Org Synth* **68** 83 1990, Evans et al. *J Am Chem Soc* **103**, 3099 1981.] **TOXIC**

**Di-*n*-butyl *n*-butylphosphonate** [78-46-6] **M 250.3, b 150-151°/10mm, 160-162°/20mm, n<sup>25</sup> 1.4302**. Purify by three recrystallisations of its compound with uranyl nitrate, from hexane. For method, see *tributyl phosphate*.

**Di-*n*-butyl cyclohexylphosphonate** [1085-92-3] **M 245.4**. The compound with uranyl nitrate is recrystallised three times from hexane. For method see *tributyl phosphate*.

**Di-*tert*-butyl dichlorosilane (DTBCl<sub>2</sub>)** [18395-90-9] **M 213.2, m -15°, b 190°/729mm, 195-197°/atm, d<sub>4</sub><sup>20</sup> 1.01**. Purify it by fractional distillation. It is a colourless liquid with a pleasant odour and does not fume in moist air, but does not titrate quantitatively with excess of dilute alkali. [Tyler et al. *J Am Chem Soc* **70** 2877 1948.]

**Di-*tert*-butylphosphine** [819-19-2] **M 146.2, b 34-35°/2mm, 38-40°/13mm, d<sub>4</sub><sup>25</sup> 0.951, pK<sub>Est</sub> ~8.7**. It is prepared by reaction of *tert*-butylmagnesium bromide with PCl<sub>3</sub> in the presence of LiAlH<sub>4</sub>, and purified by distillation in a vacuum. *Alternatively*, the intermediate *di-tert-butylchlorophosphine* [13716-10-4] is reduced separately and the product is purified by fractional distillation *in vacuo*. [Hoffmann & Schellenbeck *Chem Ber*

99 1134 1966, Hoffmann & Schellenbeck *Chem Ber* **100** 692 1967, Crofts et al. *J Chem Soc C* 331 1970.]

**Di-tert-butyl silyl bis(trifluoromethanesulfonate)** [85272-31-7] **M 440.5, b 73.5-74.5°/0.35mm, d<sub>4</sub><sup>20</sup> 1.36** (see **pK** for triflic acid). Purify it by fractional distillation at high vacuum. It is a pale yellow liquid that should be stored under argon. It is less reactive than the diisopropyl analogue. The presence of the intermediate monochloro compound can be detected by <sup>1</sup>H NMR, (CHCl<sub>3</sub>): *tert*-Bu<sub>2</sub>Si(OTf)<sub>2</sub> [δ 1.25s], but impurities have δ 1.12s for *tert*-Bu<sub>2</sub>Si(H)OTf and δ 1.19s for *tert*-Bu<sub>2</sub>HSi(Cl)OTf. [Deslongchamps *Tetrahedron Lett* **23** 4871 1982, Deslongchamps *Aldrichimica Acta* **17** 72 1984.] **TOXIC.**

**2,6-Dichlorophenol-indophenol sodium salt (2H<sub>2</sub>O)** [620-45-1] **M 326.1, ε 2.1 x 10<sup>4</sup> at 600nm and pH 8, pK<sup>30</sup> 5.7 (oxidised form), pK<sub>1</sub><sup>30</sup> 7.0, pK<sub>2</sub><sup>30</sup> 10.1 (reduced form).** Dissolve it in 0.001M phosphate buffer, pH 7.5 (alternatively, about 2g of the dye is dissolved in 80ml of M HCl), and extracted into diethyl ether. The extract is washed with water, extracted with aqueous 2% NaHCO<sub>3</sub>, and the sodium salt of the dye is precipitated by adding NaCl (30g/100ml of NaHCO<sub>3</sub> solution), then filtered off, washed with dilute NaCl solution and dried. It has λ<sub>max</sub> at 605nm. [Hiromi et al. *Anal Biochem* **101** 421 1980.] The *acetate* [24857-20-3] **M 310.1** has m 101-103° (from Et<sub>2</sub>O/petroleum ether) and 99.5-100.5° (from Et<sub>2</sub>O). [Beilstein **13** IV 1078-1079.]

**Dicyclopentylphosphine** [39864-68-1] **M 170.2, b 76-78°/0.8mm, d<sub>4</sub><sup>25</sup> 0.933, pK<sub>Est</sub> ~4.5.** Purify dicyclopentylphosphine by distillation in a vacuum in a stream of N<sub>2</sub> or Ar as it is air sensitive and must be stored in an inert atmosphere. [cf Neidergall *Chem Ber* **95** 64 1962, Beilstein **16** IV 947 for *dicyclohexylphosphine*.]

**O,O-Diethyl-S-2-diethylaminoethyl phosphorothiolate** [78-53-5] **M 269.3, m 98-99°.** Recrystallise it from isopropanol/diethyl ether. [Ailman & Magee in *Organo Phosphorus Compound* (Kosolapoff & Maier eds) Wiley Vol 7 pp 487-871 1976.]

**Diethylmethylsilane** [760-32-7] **M 102.3, b 78.4°/760mm, 77.2-77.6°/atm, d<sub>4</sub><sup>20</sup> 0.71.** Fractionally distil it through a *ca* 20-plate column, and the fraction boiling within a range of less than 0.5° is collected. It is a **flammable irritating** liquid. [Price *J Am Chem Soc* **69** 2600 1947, Beilstein **4** III 1847.]

**N,N-Diethyltrimethylsilylamine** [996-50-9] **M 145.3, b 33°/26mm, 126.8-127.1°/738mm, 126.1-126.4°, d<sub>4</sub><sup>20</sup> 0.763, n<sub>D</sub><sup>20</sup> 1.411.** Fractionate it through a 2ft vacuum-jacketed column containing Helipak packing with a reflux ratio of 10:1. [Sauer & Hasek *J Am Chem Soc* **68** 241 1946, Langer et al. *J Org Chem* **23** 50 1958, Rühlmann *J Prakt Chem* **9** 315 1959, Beilstein **4** IV 4010.]

**Diethyl trimethylsilyl phosphite** [13716-45-5] **M 210.3, b 61°/10mm, 66°/15mm, d<sub>4</sub><sup>20</sup> 0.9476, n<sub>D</sub><sup>20</sup> 1.4113.** Fractionate it under reduced pressure and has <sup>31</sup>PMR: δ<sub>P</sub> -128ppm relative to H<sub>3</sub>PO<sub>4</sub>. [Sekine et al. *J Org Chem* **46** 2097 1981, Chernyshev et al. *J Gen Chem USSR (Engl Transl)* **45** 231 1975.]

**Dihexadecyl phosphate (DHP)** [2197-63-9] **M 546.9, m 75°, pK<sub>Est</sub> ~1.2.** Recrystallise it from MeOH. Solutions of DHP were made in CHCl<sub>3</sub> or CH<sub>2</sub>Cl<sub>2</sub> for making vesicles. It has been used for making surfactant vesicles in studies of the viscosity-dependent variations of the fluorescence yield ϕ<sub>F</sub> and the polarity induced shift of the emission band maximum λ<sub>F(max)</sub> of a derivative of a [*p*-(dialkylamino)benzylidene]malononitrile [74677-08-0] fluorescence probe. [Lukac *J Am Chem Soc* **106** 4387 1984]. [Beilstein **1** IV 1880.]

**1,2-Dihydroxybenzene-3,5-disulfonic acid, di-Na salt (TIRON)** [149-45-1] **M 332.2, ε 6.9 x 10<sup>4</sup> at 260nm, pH 10.8, pK<sub>1</sub> and pK<sub>2</sub> <2 (for SO<sub>3</sub><sup>-</sup>), pK<sub>3</sub> 7.7, pK<sub>4</sub> 12.6 (for OHs of disulfonate dianion).** Recrystallise it from water [Hamaguchi et al. *Anal Chim Acta* **9** 563 1962]. It is an indicator colour reagent for Fe, Mn, Ti and Mo ions and complexes with Al, Cd, Co, Co, Fe (III), Mn, Pd, UO<sub>2</sub><sup>2+</sup>, VO<sub>2</sub><sup>2+</sup> and Zn. [Beilstein **11** IV 630.]

**(±)-Diisooctyl phenylphosphonate (bis[2-ethylhexyl] phenylphosphonate)** [49637-59-4] **M 378.5, b 204-207°/4mm, d<sub>4</sub><sup>20</sup> 0.970, n<sub>D</sub><sup>25</sup> 1.4780.** Distil it in a vacuum, percolate it through a column of alumina, then pass it through a packed column maintained at 150° to remove residual traces of volatile materials in a countercurrent

stream of N<sub>2</sub> under reduced pressure [Dobry & Keller *J Phys Chem* **61** 1448 1957]. [*Beilstein* **16** III 884.]

**Diisopropyl chlorosilane (chlorodiisopropylsilane)** [2227-29-4] **M 150.7, b 59°/8mm, 80°/10mm, 200°/738mm, d<sub>4</sub><sup>20</sup> 0.9008, n<sub>D</sub><sup>20</sup> 1.4518.** Impurities can be readily detected by <sup>1</sup>H NMR spectroscopy. Purify it by fractional distillation [Gilman & Clark *J Am Chem Soc* **69** 1499 1947, Allen et al. *J Chem Soc* 3668 1957].

**Dilongifolyl borane** [77882-24-7] **M 422.6, m 169-172°.** Wash it with dry Et<sub>2</sub>O and dry it in a vacuum under N<sub>2</sub>. It has **m 160-161°** in a sealed evacuated capillary. It is sparingly soluble in pentane, tetrahydrofuran, carbon tetrachloride, dichloromethane, and chloroform, but the suspended material is capable of causing asymmetric hydroboration. Disappearance of solid indicates that the reaction has proceeded. [Jadhav & Brown *J Org Chem* **46** 2988 1981.]

**Dimethyldichlorosilane** [75-78-5] **M 129.1, m -75.5°, b 68.5-68.7°/750mm, 70.5°/760mm, d<sub>4</sub><sup>20</sup> 1.0885, n<sub>D</sub><sup>20</sup> 1.4108.** Other impurities are chlorinated silanes and methylsilanes. Fractionate it through a 3/8in diameter 7ft Stedman column rated at 100 theoretical plates at almost total reflux. See purification of MeSiCl<sub>2</sub>. Solutions in heptane, 1,1,1-trichloroethane or 1-chloronaphthalene are used for the silanization of glassware and pipettes. [Sauer & Hadsell *J Am Chem Soc* **70** 3590 1948, *Beilstein* **4** IV 4110.]

**2,6-Dimethyl-1,10-phenanthrolinedisulfonic acid, di-Na salt (H<sub>2</sub>O) (bathocuproine-disulfonic acid di-Na salt)** [52698-84-7] **M 564.5, m ~300°, pK<sub>Est</sub> ~0 (for free acid).** Inorganic salts and some coloured species can be removed by dissolving the crude material in the minimum volume of water and precipitating by adding EtOH. The purified reagent can be obtained by careful evaporation of the filtrate. Recrystallise it from EtOH and dry it in a vacuum at room temperature in the dark. [Lorenzo et al. *J Electroanalyt Chem* **356** 43 1993, Watkins et al. *Microchem J* **16** 14 1971.]

**Dimethylphenylsilyl chloride (DMPSiCl, chlorodimethylphenylsilane, phenyl dimethyl chlorosilane)** [768-33-2] **M 170.7, b 79°/15mm, 85-87°/32mm, 196°/760mm, d<sub>4</sub><sup>20</sup> 1.017, n<sub>D</sub><sup>20</sup> 1.509.** Fractionate it through a 1.5 x 18inch column packed with stainless steel helices, or a spinning band column. [Daudt & Hyde *J Am Chem Soc* **74** 386 1952, Lewis et al. *J Am Chem Soc* **70** 1115 1948, Eaborn *J Chem Soc* 494 1953.] It is used for standardising MeLi or MeMgBr that form Me<sub>3</sub>PhSi which is estimated by GC. [Maienthal et al. *J Am Chem Soc* **76** 6392 1954, House & Respass *J Organomet Chem* **4** 95 1965, *Beilstein* **16** IV 1475.] **TOXIC and MOISTURE SENSITIVE.**

**Dimethyl thiophosphonate (dimethyl hydrogen phosphonothiolate)** [5930-72-3] **M 126.1, b 53-53.5°/16.5mm, 56-59°/9mm, 59°/16mm, d<sub>4</sub><sup>25</sup> 1.1892, n<sub>D</sub><sup>20</sup> 1.4768.** Fractionally distil the ester in a stream of dry N<sub>2</sub> or Ar at as high a vacuum as possible. Store in a sealed ampoule under Ar as it has a foul odour. The IR (film) has  $\nu_{\max}$  at ~800 cm<sup>-1</sup> (12.6 μ, P-S). [IR: McIvor et al. *Canad J Chem* **34** 1611 1956, McIvor et al. *Canad J Chem* **36** 820, 1818 1958, Tongcharoensirikul et al. *J Org Chem* **69** 2322 2004, *Beilstein* **1** IV 1258.]

**Diocetyl phenylphosphonate** [1754-47-8] **M 378.8, b 207°/4mm, d<sub>4</sub><sup>20</sup> 1.485, n<sub>D</sub><sup>25</sup> 1.4780.** Purify it as described for diisooctyl phenylphosphonate and distil under high vacuum. [*Beilstein* **16** IV 1069.]

**(1,3-Dioxalan-2-ylmethyl)triphenylphosphonium bromide** [52509-14-5] **M 429.3, m 191.5-193°, 193-195°.** Wash the crystals with Et<sub>2</sub>O, dry them in a vacuum and recrystallise them from CH<sub>2</sub>Cl<sub>2</sub>/dry Et<sub>2</sub>O to give prisnts with **m 172-174°**, which is raised to 191.5-193° on drying at 56°/0.5mm. [Cresp et al. *J Chem Soc, Perkin Trans I* 37 1974.]

**Diphenyldiselenide** [1666-13-3] **M 312.1, m 62-64.** Crystallise it twice from hexane [Kice & Purkiss *J Org Chem* **52** 3448 1987]. [*Beilstein* **6** IV 1781.]

**Diphenyl hydrogen phosphate (diphenyl phosphate)** [838-85-7] **M 250.2, m 99.5°, pK<sup>20</sup> 0.26.** Crystallise it from CHCl<sub>3</sub>/petroleum ether. [Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973, *Beilstein* 6 IV 714.]

**4,7-Diphenyl-1,10-phenanthrolinedisulfonic acid, di-Na salt 3H<sub>2</sub>O (bathophenanthroline-disulfonic acid di-Na salt)** [52746-49-3] **M 590.6, m 300°, pK<sub>Est</sub>~0 (for free acid).** Dissolve crude sample in the minimum volume of water and add EtOH to precipitate the contaminants. Carefully evaporate the filtrate to obtain pure material. It forms a dark red complex with Fe<sup>2+</sup> with  $\lambda_{\max}$  535nm ( $\epsilon$  2.23 x 10<sup>4</sup>mol<sup>-1</sup> cm<sup>-1</sup>) [Imasaka et al. *Anal Chim Acta* 115 407 1980]. It is prepared by sulfonating bathophenanthroline with ClSO<sub>3</sub>H: to 100g of bathophenanthroline is added 0.5ml of Fe free ClSO<sub>3</sub>H and heated over a flame for 30 seconds (**CARE**). Cool and carefully add 10ml of pure distilled H<sub>2</sub>O and warm on a water bath with stirring till all solid dissolves. A stock solution is made by diluting 3ml of this reagent to 100ml with 45% aqueous NaOAc, filter off the solid and store in a dark bottle. In this way it is stable for several months. [Landers & Bennie *Am J Clinical Pathology* 29 590 1958.]

**Diphenylphosphinic acid** [1707-03-5] **M 218.2, m 194-195°, pK<sup>20</sup> 1.72.** Recrystallise it from 95% EtOH and dry it under vacuum at room temperature. [see Kosolapoff *Organophosphorus Compounds* J Wiley, NY, 1950, Kosolapoff and Maier *Organic Phosphorus Compounds* Wiley-Interscience, NY, 1972-1976, *Beilstein* 16 IV 1036.]

**Diphenyl phosphorazidate (DPPA, diphenyl phosphoryl azide, phosphoric acid diphenyl ester azide)** [26386-88-9] **M 275.2, b 134-136°/0.2mm, 137°/0.17mm, d<sub>4</sub><sup>25</sup> 1.277, n<sub>D</sub><sup>20</sup> 1.551.** This azide and nitrene source is prepared by stirring a mixture of diphenyl phosphorochloridate (56.8g, 210mmol, freshly distilled at 165-168°/5mm, see [2524-64-3]), sodium azide (16.3g, 250mmol) and anhydrous Me<sub>2</sub>CO (300ml, dried over K<sub>2</sub>CO<sub>3</sub> and freshly distilled) in a dry atmosphere at 20-25° for 21 hours. The *lachrymatory* mixture is filtered in a fume cupboard, and the filtrate is concentrated in vacuum. The residue is purified by distillation through a short Vigreux column keeping the bath temperature below 200° to avoid decomposition of the *azidate*, and collecting the fraction with b 134-136°/0.2mm (49-52g, 84-89%). It is a colourless non-explosive oil which can be stored for long periods if it is protected from moisture and light. It has IR (neat) with  $\nu_{\max}$  at 3060 (w, C-H), 2170 (s, -N<sub>3</sub>), 1590 (m), 1490 (s, arene C=C), 1270 (m, P=O) and 960 (s, P-O-aryl) cm<sup>-1</sup>; and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  at 7.0-7.3 (br s, aryl-H). [Shioiri & Yamada *Org Synth Coll Vol VII* 206 1990.]

It is a useful activating agent and a versatile reagent in organic synthesis (see references cited in *Org Synth Coll Vol VII* 206 1990 above). It is commercially available on a 100-200 mesh, 1% cross-linked polydivinylbenzene support and bound (*via para* linkage) through one of its phenyl rings (PS-DPPA, 1.0-1.5mmol/g N<sub>3</sub> loading).

**Diphenyl phosphoro chloridate (diphenyl phosphoryl chloride)** [2524-64-3] **M 268.6, b 141°/1mm, 194°/13mm, 314-316°/272mm, d<sub>4</sub><sup>30</sup> 1.2960, n<sub>D</sub><sup>35</sup> 1.5490.** Fractionally distil it in a good vacuum; better use a spinning band column. [Walsh *J Am Chem Soc* 81 3023 1959, IR: Bellamy & Beecher *J Chem Soc* 475 1952, *Beilstein* 6 IV 737.]

**Diphenylsilane** [775-12-2] **M 184.3, b 75-76°/0.5mm, 113-114°/9mm, 124-126°/11mm, 134-135°/16mm, d<sub>4</sub><sup>20</sup> 1.0027, n<sub>D</sub><sup>20</sup> 1.5802.** Dissolve it in Et<sub>2</sub>O, mix slowly with ice-cold 10% AcOH. The Et<sub>2</sub>O layer is then shaken with H<sub>2</sub>O until the washings are neutral to litmus. Dry over Na<sub>2</sub>SO<sub>4</sub>, evaporate the Et<sub>2</sub>O and distil the residual oil under reduced pressure using a Claisen flask with the take-off head modified into a short column. Ph<sub>2</sub>SiH<sub>2</sub> boils at 257°/760mm, but it cannot be distilled at this temperature because exposure to air leads to flashing, decomposition and formation of silica. It is a colourless, odourless oil, miscible with organic solvents but not H<sub>2</sub>O. A possible impurity is Ph<sub>3</sub>SiH which has **m** 43-45° and would be found in the residue. [West & Rochow *J Org Chem* 18 303 1953, Benknesser et al. *J Am Chem Soc* 74 648 1952, Gilman & Zuech *J Am Chem Soc* 81 5925 1959, *Beilstein* 16 IV 1366.]

**Diphenylsilanediol** [947-42-2] **M 216.3, m 148°(dec).** Recrystallise it from CHCl<sub>3</sub>/methyl ethyl ketone. [*Beilstein* 16 IV 1523.]

**Diphenyl *p*-tolyl phosphate** [26444-49-5] **M 340.3, m 18-29°, b 156-158°/0.002mm, n<sub>D</sub><sup>25</sup> 1.5758.** Distil it in a vacuum, then percolate it through a column of alumina. Finally, pass it through a packed column maintained at 150° to remove traces of volatile impurities in a countercurrent stream of nitrogen under reduced pressure. [Dobry & Keller *J Phys Chem* **61** 1448 1947, *Beilstein* **6** IV 2130.]

**Disodium 4,5(1,8)-dihydroxynaphthalene-2,7(3,6)-disulfonate 2H<sub>2</sub>O (Chromotropic acid di-Na salt)** [5808-22-0] **M 400.3, m >300°, pK<sub>1</sub> 0.61(SO<sub>3</sub><sup>-</sup>), pK<sub>2</sub> 0.7(SO<sub>3</sub><sup>-</sup>), pK<sub>3</sub> 5.45(OH), pK<sub>4</sub> 15.5(OH).** Recrystallise it from H<sub>2</sub>O or H<sub>2</sub>O by adding EtOH. It complexes with Ag, ClO<sub>3</sub><sup>-</sup>, Cr, Hg, NO<sub>2</sub><sup>-</sup>, NO<sub>3</sub><sup>-</sup> and Ti. [*Beilstein* **11** H 72, **11** I 307, **11** II 174, **11** III 174, **11** IV 576.]

**Disodium ethylenebis(dithiocarbamate) (Nabam)** [142-59-6] **M 436.5, pK<sub>Est</sub> ~3.0.** It crystallises (as *hexahydrate*) from aqueous ethanol. It is a skin irritant. [*Beilstein* **4** III 149, **4** IV 234.]

**Disodium-β-glycerophosphate** [819-83-0 (4H<sub>2</sub>O), 13408-09-8 (5H<sub>2</sub>O)] **M 216.0, m 102-104°, pK<sub>2</sub><sup>25</sup> 6.66 (free acid).** Crystallise it from water. [Attwood et al. *Biochem J* **253** 389 1988, *Beilstein* **1** IV 2766.]

**Disodium naphthalene-1,5-disulfonate** [1655-29-4] **M 332.3, pK<sub>Est</sub> ~0.** Recrystallise it from aqueous acetone. It darkens at high temperatures. [Okahata et al. *J Am Chem Soc* **108** 2863 1986, *Beilstein* **11** IV 561.]

**Disodium 4-nitrophenylphosphate (6H<sub>2</sub>O)** [4264-83-9] **M 371.1.** Dissolve it in hot aqueous MeOH, filter and precipitate it by adding Me<sub>2</sub>CO. Wash the solid with Me<sub>2</sub>CO and repeat the purification. Aqueous MeOH and Et<sub>2</sub>O can also be used as solvents. The white fibrous crystals contain less than 1% of free 4-nitrophenol [assay: Axelrod *J Biol Chem* **167** 57 1947]. [*Beilstein* **6** IV 1327.]

**Disodium phenylphosphate (2H<sub>2</sub>O)** [3279-54-7, 66778-08-3 (2H<sub>2</sub>O)] **M 254.1, pK<sub>1</sub><sup>25</sup> 1.46, pK<sub>2</sub><sup>25</sup> 6.29 [for PhPO(OH)<sub>2</sub>].** Dissolve it in a minimum amount of methanol, filtering off any insoluble residue of inorganic phosphate, then precipitate it by adding an equal volume of Et<sub>2</sub>O. Wash the solid with Et<sub>2</sub>O and dry it in a vacuum [Tsuboi *Biochim Biophys Acta* **8** 173 1952]. [*Beilstein* **6** IV 708.]

**Di-*p*-tolyl phenylphosphonate** [94548-75-1] **M 388.3, n<sub>D</sub><sup>25</sup> 1.5758.** Purify as described under diisooctyl phenylphosphonate.

**1,3-Divinyl-1,1,3,3-tetramethyldisiloxane** [2627-95-4] **M 186.4, m -99.7°, b 128-129°/atm, 139°/760mm, d<sub>4</sub><sup>20</sup> 0.811, n<sub>D</sub><sup>20</sup> 1.4122.** Dissolve it in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dry it over CaCl<sub>2</sub>, evaporate the solvent and distil the residue. It is an IRRITANT. [Kantor et al. *J Am Chem Soc* **77** 1685 1955, Bazant & Matousek *Col Czech Chem Comm* **24** 3758 1959, *Beilstein* **4** IV 4080.]

**Eriochrome Black T [3-hydroxy-7-nitro-4-(1-hydroxy-2-naphthylazo)naphthalene-1-sulfonic acid Na salt]** [1787-61-7] **M 416.4. A<sub>1cm</sub><sup>%</sup> (λ<sub>max</sub>) 656(620nm) at pH 10, using the dimethylammonium salt, pK<sub>2</sub><sup>25</sup> 5.81, pK<sub>3</sub><sup>25</sup> 11.55.** The sodium salt (200g) is converted to the free acid by stirring with 500ml of 1.5M HCl, and, after several minutes, the slurry is filtered on a sintered-glass funnel. The process is repeated and the material is air dried after washing with acid. It is then extracted with \*benzene for 12 hours in a Soxhlet extractor, the \*benzene solution is evaporated and the residue is air dried. A further desalting with 1.5M HCl (1L) is followed by crystallisation from dimethylformamide (in which it is very soluble) by forming a saturated solution at the boiling point, and allowing to cool slowly. The crystalline dimethylammonium salt so obtained is washed with \*benzene and treated repeatedly with dilute HCl to give the insoluble *free acid* which, after air drying, is dissolved in alcohol, filtered and evaporated. The final material is air dried, then dried in a vacuum desiccator over Mg(ClO<sub>4</sub>)<sub>2</sub>. The purified acid is converted to the dimethylammonium salt with Me<sub>2</sub>NH. [Diehl & Lindstrom *Anal Chem* **31** 414 1959]. It is an indicator in the complexometry of alkaline earth metals. [*Beilstein* **16** IV 429.]

**Eriochrome Blue Black B (Mordant Black, 3, 3-hydroxy-4-(1-hydroxy-2-naphthylazo)-naphthalene-1-**

**sulfonic acid Na salt** [3564-14-5] M 416.4,  $pK_2^{25}$  7.0,  $pK_3^{25}$  13.5. Free it from metallic impurities by three precipitations from aqueous solution by addition of HCl. The precipitated dye is dried at 60° under vacuum and converted into the Na salt with the calculated amount of alkali. It is an indicator in the complexometry of Al, Fe and Zr. [Beilstein 16 III 319.]

**Eriochrome Blue Black R (Palatine Chrome Black 6BN, Mordant Black 17, 3-hydroxy-4-(2-hydroxy-1-naphthylazo)naphthalene-1-sulfonic acid Na salt** [2538-85-4] M 416.4,  $pK_2^{25}$  7.0,  $pK_3^{25}$  13.5. Free it from metallic impurities by three precipitations from aqueous solution by addition of HCl. The precipitated dye is dried at 60° under vacuum and converted into the Na salt with the calculated amount of alkali. It is an indicator in the complexometry of Al, Fe and Zr. [Beilstein 16 H 297, 16 IV 428.]

**Ethoxycarbonylmethylene triphenylphosphonium bromide** [1530-45-6] M 429.3, m 155-155.5°, 158°(dec). Wash it with petroleum ether (b 40-50°) and recrystallise it from  $CHCl_3/Et_2O$  and dry it in a high vacuum at 65°. [Isler et al. *Helv Chim Acta* 40 1242 1957, Wittig & Haag *Chem Ber* 88 1654, 1664 1955.]

**(Ethoxycarbonylmethylene)triphenylphosphorane [ethyl (triphenylphosphoranylidene)-acetate]** [1099-45-2] M 348.4, m 116-117°, 128-130°. Crystallise it by dissolving it in AcOH and adding petroleum ether (b 40-50°) to give colourless plates. Its UV has  $\lambda_{max}$  ( $A_{1\%}^{1\text{cm}}$ ) at 222nm (865) and 268nm (116) [Isler et al. *Helv Chim Acta* 40 1242 1957, Beilstein 16 IV 977.]

**Ethylarsonic acid** [507-32-4] M 154.0, m 94-95°, 99.5°,  $pK_1$  4.72 ( $As(OH)O^-$ ),  $pK_2$  8.00 [ $AsO_2^{2-}$ ]. Crystallise it from ethanol. Also it dissolves in excess EtOH, filter off the insoluble matter (salts?), evaporate the filtrate to dryness and recrystallise the residue from small volumes of EtOH or  $H_2O$ . [Quick & Adams *J Am Chem Soc* 44 805 1922, Banks et al. *J Am Chem Soc* 69 927 1947, Beilstein 4 H 614, 4 II 997, 4 III 1823, 4 IV 3682.]

**Ethyl Orange (sodium 4,4'-diethylaminophenylazobenzenesulfonate)** [62758-12-7] M 355.4,  $pK_{Est} \sim 3.8$ . Recrystallise it twice from water. [Beilstein 16 IV 511.]

**Ethyl trimethylsilylacetate** [4071-88-9] M 160.3, b 75.5°/42mm, 157°/730mm,  $d_4^{20}$  0.8762,  $n_D^{20}$  1.4149. Purify it by distilling ca 10g of reagent through a 15cm, Vigreux column and then redistilling it through a 21cm glass helices-packed column [Hauze & Hauser *J Am Chem Soc* 75 994 1953]. Alternatively, dissolve it in  $Et_2O$ , wash with  $H_2O$ , dilute  $Na_2CO_3$ , dry over  $Na_2CO_3$ , evaporate  $Et_2O$ , and distil it through a column of 15 theoretical plates. [Gold et al. *J Am Chem Soc* 70 2874 1948, Beilstein 4 IV 3974.]

**Ethyl 3-(trimethylsilyl)propionate** [17728-88-0] M 174.3, b 93°/40mm, 178°-180°/atm,  $d_4^{20}$  0.8763,  $n_D^{20}$  1.4198. Dissolve it in  $Et_2O$ , wash this with  $H_2O$ , dilute  $Na_2CO_3$ , dry ( $Na_2SO_4$ ), evaporate  $Et_2O$  and fractionally distil. [Sommer & Marans *J Am Chem Soc* 72 1935 1950, Beilstein 4 IV 3975.]

**Ethyl triphenylphosphonium bromide** [1530-32-1] M 371.3, m 203-205°. Recrystallise it from  $H_2O$  and dry it in high vacuum at 100°. Its IR has bands at 1449, 1431 and 997 $cm^{-1}$ . [Wittig & Wittenberg *Justus Liebigs Ann Chem* 606 1 1957, Bergmann & Dusza *J Org Chem* 23 1245 1958, Beilstein 16 IV 982.]

**Ethynyl trimethylsilane** [1066-54-2] M 98.2, b 53°/atm, 52.5°/atm,  $d_4^{20}$  0.71,  $n_D^{20}$  1.3871. Distil it through an efficient column. The IR has bands at  $\nu_{max}$  2041 ( $C\equiv C$ ) and 3289 ( $\equiv C-H$ )  $cm^{-1}$ . [Kröhnke & Goss *Chem Ber* 92 30 1959, Beilstein 4 IV 3937.]

**Fluorotrimethylsilane (trimethylsilyl fluoride, TMSF)** [420-56-4] M 92.2, m -74°, b 16°/760mm, 19°/730mm,  $d^0$  0.793. It is a FLAMMABLE gas which is purified by fractional distillation through a column at low temperature and with the exclusion of air. Preferably work in a vacuum line. [Booth & Suttle *J Am Chem Soc* 68 2658 1946, Reid & Wilkins *J Chem Soc* 4029 1955, Beilstein 4 IV 4007.]

**Hexaethylidisiloxane** [924-49-0] M 246.5, b 114-115°/16mm, 235.5°/760mm,  $d_4^{20}$  0.8443,  $n_D^{20}$  1.4330. Distil in a vacuum, but it can be distilled at atmospheric pressure without decomposition. It is characterised by completely dissolving in conc H<sub>2</sub>SO<sub>4</sub>. [Eaborn *J Chem Soc* 3077 1950, *Beilstein* 4 IV 4055.]

**2,2,4,4,6,6-Hexamethylcyclotrisiloxane** [1009-93-4] M 219.5, m -10°, b 81-82°/19mm, 111-112°/85mm, 188°/756mm,  $d_4^{20}$  0.9196,  $n_D^{20}$  1.448. Purify it by fractional distillation at atmospheric pressure until the temperature reaches 200°. It is *moisture sensitive*. The residue in the flask is mostly octamethylcyclotetrasilazane. [Brewer & Haber *J Am Chem Soc* 70 3888 1948, *Beilstein* 4 III 1887.]

**Hexamethyldisilane** [1450-14-2] M 164.4, m 9-12°, b 113.1°/750mm,  $d_4^{20}$  0.7272,  $n_D^{20}$  1.4229. The most likely impurity is trimethylchlorosilane (*cf* boiling point). Wash it with H<sub>2</sub>O, cold conc H<sub>2</sub>SO<sub>4</sub>, H<sub>2</sub>O again, then aqueous NaHCO<sub>3</sub>, dry over CaSO<sub>4</sub> and fractionate at atmospheric pressure. [Brown & Fowles *J Chem Soc* 2811 1958.] A grossly impure sample (25% impurities) was purified by repeated spinning band distillation. This lowered the impurity level to 500 ppm. The main impurity was identified as 1-hydroxypentamethyldisilane. [*Beilstein* 4 IV 4277.]

**Hexamethyldisilazane [HMDS, (Me<sub>3</sub>Si)<sub>2</sub>NH]** [999-97-3] M 161.4, b 125-125.6°/atm, 126°/760mm,  $d_4^{20}$  0.7747,  $n_D^{20}$  1.407. A possible impurity is Me<sub>3</sub>SiCl. Wash it well with petroleum ether and fractionate it through a vacuum jacketed column packed with Helipac using a reflux ratio of 10:1. [Langer et al. *J Org Chem* 23 50 1958, *Beilstein* 4 IV 4014.] It forms a lithium salt, (Me<sub>3</sub>Si)<sub>2</sub>NLi [4039-32-1], which is soluble in THF and hexane; and ~ 1M solutions in these solvents are available commercially.

**Hexamethyldisiloxane (HMDSO)** [107-46-0] M 162.4, m -59°, b 99.4°/760mm, 100.4°/764mm,  $d_4^{20}$  0.7633,  $n_D^{20}$  1.3777. Fractionally distil through a column packed with glass helices with *ca* 15 theoretical plates. It is highly **flammable** and is an **irritant**. [Mills & McKenzie *J Am Chem Soc* 76 2672 1954, Csakvari et al. *J Organometal Chem* 107 287 1976, *Beilstein* 4 IV 4018.]

**Hexamethylphosphoric triamide (HMPA)** [680-31-9] M 179.2, f 7.2°, b 68-70°/1mm, 235°/760mm,  $d_4^{20}$  1.024,  $n_D^{20}$  1.460. The industrial synthesis is usually by treatment of POCl<sub>3</sub> with excess of dimethylamine in isopropyl ether. Impurities are water, dimethylamine and its hydrochloride. It is purified by refluxing over BaO or CaO at about 4mm pressure in an atmosphere of nitrogen for several hours, then distilled from sodium at the same pressure. The middle fraction (b *ca* 90°) is collected, refluxed over sodium under reduced pressure under nitrogen and distilled. It is kept in the dark under nitrogen, and stored in solid CO<sub>2</sub>. It can also be stored over 4A molecular sieves.

*Alternatively*, it is distilled under vacuum from CaH<sub>2</sub> at 60° and is crystallised twice in a cold room at 0°, seeding the liquid with crystals obtained by cooling in liquid nitrogen. After about two-thirds are frozen, the remaining liquid is drained off [Fujinaga et al. *Pure Appl Chem* 44 117 1975]. For tests of purity see Fujinaga et al. in *Purification of Solvents*, Coetzee Ed., Pergamon Press, Oxford, 1982. For efficiency of desiccants in drying HMPA see Burfield and Smithers [*J Org Chem* 43 3966 1978, and Sammes et al. *J Chem Soc, Faraday Trans 1* 281 1986]. [*Beilstein* 4 IV 284.] **CARCINOGEN**.

**Hexamethylphosphorous triamide [HMPT, tris(dimethylamino)phosphine]** [1608-26-0] M 163.2, m 7.2°, b 49-51°/12mm, 162-164°/12mm,  $d_4^{20}$  0.989,  $n_D^{20}$  1.466. It may contain more than 1% of phosphoric triamide. The yellow oil is first distilled at atmospheric pressure, then under reduced pressure and stored under N<sub>2</sub>. It is air sensitive, **TOXIC**, and *should not be inhaled*. It is absorbed through the skin. [Mark *Org Synth Coll Vol V* 602 1973, *Beilstein* 4 IV 274.]

**Hydroquinone-2-sulfonic acid K salt** [21799-87-1] M 228.3, m 250°(dec),  $pK_{Est(1)} \sim 1$ ,  $pK_{Est(2)} \sim 8.5$ ,  $pK_{Est(3)} \sim 11$ . Recrystallise it from water or EtOH. [*Beilstein* 11 I 70, 11 II 170, 11 III 570.]

**Hydroxynaphthol Blue tri-Na salt** [63451-35-4] M 620.5, m dec on heating,  $pK_{Est} < 0$ . The crude material is treated with hot EtOH to remove soluble impurities, then dissolve in 20% aqueous MeOH and chromatographed on a cellulose powder column with propanol/EtOH/water (5:5:4) as eluent. The upper of three

zones are eluted to give the pure dye which is precipitated as the *monosodium salt trihydrate* by adding conc HCl to the concentrated eluate [Ito & Ueno *Analyst* **95** 583 1970]. It can be converted to the trisodium salt with the calculated amount of alkali.

**Indigocarmine** (2[1,3-dihydro-3-oxo-5-sulfo-2H-indol-2-ylidene]-2,3-dihydro-3-oxo-1H-indole-5-sulfonic acid di-Na salt) [860-22-0] **M 466.4**, **pK<sub>1</sub><sup>20</sup> 2.8**, **pK<sub>2</sub><sup>20</sup> 12.3**. Its solubility in H<sub>2</sub>O is 1g/100ml at 25°. It has been purified by dissolving in H<sub>2</sub>O, filtering and adding EtOH to cause the salt to separate. Wash the solid with EtOH, Et<sub>2</sub>O and dry *in vacuo*. [Vörlander & Schubert *Chem Ber* **34** 1860 1901, UV: Smit et al. *Anal Chem* **27** 1159 1955, Preisler et al. *J Am Chem Soc* **81** 1991 1959, *Beilstein* **25** IV 1975.]

**Iodomethyl trimethylsilane** [4206-67-1] **M 214.1**, **b 139.5°/744mm**, **d<sub>4</sub><sup>20</sup> 1.44**, **n<sub>D</sub><sup>25</sup> 1.4917**. If slightly violet in colour, wash it with aqueous 1% sodium metabisulfite, H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>) it and fractionally distil it at 760mm. [Whitmore & Sommer *J Am Chem Soc* **68** 481 1946, *Beilstein* **4** IV 3878.]

**Iodotrimethylsilane (trimethylsilyl iodide, TMSI)** [16029-98-4] **M 200.1**, **b 106.8°/742mm**, **107.5°/760mm**, **d<sub>4</sub><sup>20</sup> 1.470**. Add a little antimony powder and fractionate with this powder in the still. Stabilise the distillate with 1% wt of Cu powder. [Eaborn *J Chem Soc* 3077 1950, *Beilstein* **4** IV 4009.]

**Isopropylidimethyl chlorosilane** [3634-56-8] **M 140.7**, **b 109.8-110.0°/738mm**, **d<sub>4</sub><sup>20</sup> 0.88**, **n<sub>D</sub><sup>20</sup> 1.4158**. Probable impurity is Me<sub>3</sub>SiCl (**b 56.9°/783mm**) which can be removed by efficient fractional distillation. [Sommer et al. *J Am Chem Soc* **76** 801 1954, *Beilstein* **4** IV 4067.]

**N-Lauroyl-N-methyltaurine sodium salt (sodium N-decanoyl-N-methyl-2-aminoethane sulfonate)** [4337-75-1] **M 343.5**, **pK<sub>Est</sub> ~1.5**. It is prepared from methyldecanoate (at 180° under N<sub>2</sub>) or decanoyl chloride and sodium N-methylethane sulfonate and purified by dissolving it in H<sub>2</sub>O and precipitating by addition of Et<sub>2</sub>O. It decomposes on heating. [Desseigne & Mathian *Mém Services Chim Etat Paris* **31** 359 1944, cf. *Chem Abstr* **41** 705 1947.]

**Lawesson's Reagent** [LR, 2,4-bis(4-methoxyphenyl)1,3,2,4-dithiadiphosphetane 2,4-disulfide, *p*-methoxyphenylthiophosphonic acid cyclic di(thioanhydride)] [19172-47-5] **M 404.5**, **m 228-229.5° (sintering at 215°), 228-230°**. The reagent has been washed with anisole or C<sub>6</sub>H<sub>6</sub>\* and dried in a vacuum over paraffin wax and P<sub>2</sub>O<sub>5</sub>, and heated above room temperature. It does not lose the solvents completely (possibly due to formation of clathrates). It loses anisole on heating at ~230°. Analytically pure reagent was obtained by recrystallisation from *o*-dichlorobenzene, washing the crystals with C<sub>6</sub>H<sub>6</sub>\* and drying *in vacuo* at 150°. It is extremely moisture sensitive, liberating H<sub>2</sub>S, and should be stored in sealed containers, and preferably weighed in a dry-box. Its molecular weight is consistent with the formula (C<sub>7</sub>H<sub>7</sub>OPS<sub>2</sub>)<sub>2</sub> but slowly polymerises in solution at ~80-85°. Polymerisation is more rapid in polar than non-polar solvents. It is a very useful thiating reagent particularly for converting hydroxy to thiol substituents. [Nishio *J Chem Soc, Perkin Trans I* 1113 1993, *Tetrahedron Lett* **47** 9329 2006, *Beilstein* **16** IV 1113.]

**Lead diethyldithiocarbamate** [17549-30-3] **M 503.7**, **pK<sub>1</sub><sup>25</sup> 3.36 (for N,N-diethyldithio-carbamate)**. Wash it with H<sub>2</sub>O and dry it at 60-70°, or dissolve it in the minimum volume of CHCl<sub>3</sub> and add the same volume of EtOH. Collect the solid that separates and dry it as before. *Alternatively*, recrystallise it by slow evaporation of a CHCl<sub>3</sub> solution at 70-80°. Filter the crystals, wash them with H<sub>2</sub>O until all Pb<sup>2+</sup> ions are eluted (check by adding chromate) and then dry it at 60-70° for at least 10 hours. [Lo et al. *Analyt Chem* **49** 1146 1977.]

**Lissamine Green B {1-[bis-(4,4'-dimethylaminophenyl)methyl]-2-hydroxynaphthalene-3,6-disulfonic acid sodium salt, Acid Green 50}** [3087-16-9] **M 576.6**, **m >200°(dec)**, **CI 44090**, **λ<sub>max</sub> 633nm**. Crystallise it from EtOH/water (1:1, v/v). **Irritant**. [*Beilstein* **14** II 574.]

**Lissapal C (mainly sodium salt of 9-octadecene-1-sulfate)** [2425-51-6]. Reflux the salt with 95% EtOH, then

filter to remove insoluble inorganic electrolytes. The alcoholic solution is then concentrated, and the residue is poured into dry acetone. The precipitate is filtered off, washed in acetone and dried under vacuum [Biswas & Mukerji *J Phys Chem* **64** 1 1960].

**Lissapol LS (mainly sodium salt of anisidine sulfate)** [28903-20-0]. Reflux the salt with 95% EtOH, then filter to remove insoluble inorganic electrolytes. The alcoholic solution is then concentrated, and the residue is poured into dry acetone. The precipitate is filtered off, washed in acetone and dried under vacuum [Biswas & Mukerji *J Phys Chem* **64** 1 1960].

**Lithium dodecylsulfate** [2044-56-6] **M 272.3**. Recrystallise this detergent twice from absolute EtOH and dry it under vacuum. Critical Micellar Concentration (CMC) in H<sub>2</sub>O is 8.77 x 10<sup>-3</sup>M. [Mukerjee et al. *J Phys Chem* **71** 4166 1967, Mysels & Dulin *J Colloid Sci* **10** 461 1955, *Beilstein* **1** IV 1847.]

**Lithium trimethylsilanolate (trimethylsilanol Li salt)** [2004-14-0] **M 96.1, m 120°(dec in air)**. Wash it with Et<sub>2</sub>O and petroleum ether. It sublimes at 180°/1mm as fine transparent needles. [Tatlock & Rochnow *J Org Chem* **17** 1555 1952, *Beilstein* **4** IV 3992.] Suspected **CARCINOGEN**.

**Magnesium dodecylsulfate** [3097-08-3] **M 555.1**. Recrystallise it three times from EtOH and dry it in a vacuum. [*Beilstein* **1** I 1788, **1** IV 1849.]

**Magnesium trifluoromethanesulfonate** [60871-83-2] **M 322.4, m >300°**. Wash it with CH<sub>2</sub>Cl<sub>2</sub> and dry it at 125°/2 hour and 3mmHg. [Corey & Shimoji *Tetrahedron Lett* **24** 171 1983, *Beilstein* **3** IV 34.]

**Manganous ethylenebis(dithiocarbamate) (Maneb)** [12427-38-2] **M 265.3, pK<sub>Est</sub> ~3.0 (for —NCSSH)**. Crystallise this fungicide from EtOH. It is soluble in CHCl<sub>3</sub>. It is a skin irritant. [*Beilstein* **4** III 149, **4** IV 234.]

**1R,2S,5R-Menthyl phosphorochloridite (MenOPCl<sub>2</sub>)** [95456-31-8] **M 257.1, b 110-112°/2.0mm, (62°/150mm ? also reported)**. MenOPCl<sub>2</sub> is prepared by adding a solution (1R,2S,5R)(-)-menthol (78g, 0.50mol) in CH<sub>2</sub>Cl<sub>2</sub> (50ml, or THF) dropwise to a solution of PCl<sub>3</sub> (137.5g, 1.0mol) in CH<sub>2</sub>Cl<sub>2</sub> (50ml, or THF) during 30 minutes. The mixture is stirred for 1 hour under a N<sub>2</sub> atmosphere, the volatiles are removed at ~25°/100mm, pure MenOPCl<sub>2</sub> distilled at ~62°/150mm quantitatively. It has <sup>1</sup>H NMR (400MHz, C<sub>6</sub>D<sub>6</sub>, TMS) with δ at 0.64 (m, 1H, H-4), 0.77 (m, 10H, H-3, H-8, H-9, H-10), 1.15 (M, 3H, H-2, H-5, H-6), 1.44 (m, 2H, H-3, H-4), 1.98 (M, 1H, H-7), 2.30 (M, 1H, H-6), 4.40 (m, 1H, H-1); the <sup>13</sup>C{<sup>1</sup>H} NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, TMS) with δ at 16.1 (s, C-8), 21.2 (s, C-9), 22.0 (s, C-10), 22.8 (s, C-3), 25.1 (s, C-7), 31.4 (s, C-5), 33.8 (s, C-4), 43.3 (s, C-6), 48.7 (s, C-2), 83.5 (d, <sup>2</sup>J<sub>PC</sub> = 9.8Hz, C-1); and the <sup>31</sup>P NMR (121.4MHz, THF, external ref 85% H<sub>3</sub>PO<sub>4</sub>) with δ at 175.6 or (at 162MHz, C<sub>6</sub>D<sub>6</sub>) with δ at 176.1 (d, <sup>3</sup>J<sub>PH</sub> = 13.5Hz). [Alexander et al. *Organometallics* **19** 2700 2000, Brunel & Buono *J Org Chem* **58** 7313 1993, Totland et al. *Macromolecules* **29** 6114 1996]. Note that Hey-Hawkins and co-workers found that a 5 molar excess of PCl<sub>3</sub> in a solvent-free system gave 98% of dichloridite; but when a 1:1 ratio of menthol to PCl<sub>3</sub> was used a mixture of mono- and di- menthyl phosphorochloridites were obtained [Hey-Hawkins et al. *Eur J Inorg Chem* 2776 2009].

**2-Mercaptopyridine N-oxide sodium salt (pyridinethione or pyrithione sodium salt)** [3811-73-2] **M 149.1, m ~250°(dec), pK<sub>1</sub> -1.95, pK<sub>2</sub> 4.65**. When recrystallised from water, it assayed as 98.7% based on AgNO<sub>3</sub> titration [Krivis et al. *Anal Chem* **35** 966 1963; see also Krivis et al. *Anal Chem* **48** 1001 1976, and Barton & Crich *J Chem Soc, Perkin Trans 1* 1603, 1613 1986]. [*Beilstein* **21**/7 V 151.]

**Metanil Yellow (3{4-phenylamino}phenylazo)-benzenesulfonic acid)** [587-98-4] **M 375.4, pK<sub>Est</sub> <0**. It can be salted out from water three times with sodium acetate, then repeatedly extracted with EtOH [McGrew & Schneider *J Am Chem Soc* **72** 2547 1950]. [*Beilstein* **16** II 168.]

**(Methoxycarbonylmethyl)triphenylphosphorane [methyl (triphenylphosphoranylidene)-acetate]** [2605-67-6] **M 334.4, m 162-163°, 169-171°**. Crystallise it by dissolving in it AcOH and adding petroleum ether (b

40-50°) to give colourless plates. Its UV has  $\lambda_{\max}$  ( $A_{1\text{mm}}^{1\%}$ ) at 222nm (865) and 268nm (116) [Isler et al. *Helv Chim Acta* **40** 1242 1957]. [Beilstein **16** IV 977.]

**Methoxycarbonylmethyltriphenylphosphonium bromide** [1779-58-4] **M 415.3, m 163°, 165-170°(dec)**. Wash it with petroleum ether (b 40-50°), recrystallise it from  $\text{CHCl}_3/\text{Et}_2\text{O}$  and dry it at high vacuum at 65°. [Isler et al. *Helv Chim Acta* **40** 1242 1957, Wittig & Haag *Chem Ber* **88** 1654, 1664 1955, Beilstein **16** IV 981.]

**Methoxymethyl trimethylsilane (trimethylsilylmethyl methyl ether)** [14704-14-4] **M 118.3, b 83°/740mm,  $d_4^{25}$  0.758,  $n_D^{25}$  1.3878**. It forms an azeotrope with MeOH (b 60°). If it contains MeOH (check IR for bands above  $3000\text{cm}^{-1}$ ), then wash with  $\text{H}_2\text{O}$  and fractionate. A possible impurity could be chloromethyl trimethylsilane (b 97°/740mm). [Speier *J Am Chem Soc* **70** 4142 1948, Beilstein **4** III 1844.] It is an IRRITATING FLAMMABLE liquid.

**1-Methoxy-2-methyl-1-trimethylsiloxypropene (dimethyl ketene methyl trimethylsilyl acetal)** [31469-15-5] **M 174.3, b 121-122°/0.35mm, 125-126°/0.4mm, 148-150°/atm,  $d_4^{20}$  0.86**. Add  $\text{Et}_2\text{O}$ , wash with cold  $\text{H}_2\text{O}$ , dry ( $\text{Na}_2\text{SO}_4$ ), filter, evaporate  $\text{Et}_2\text{O}$ , and distil the oily residue in a vacuum. [Ainsworth et al. *J Organometal Chem* **46** 59 1972.]

**trans-1-Methoxy-3-(trimethylsilyloxy)-1,3-butadiene (Danishefsky's diene)** [54125-02-9] **M 172.3, b 68-69°/14mm, 70-72°/16mm,  $d_4^{20}$  0.885,  $n_D^{20}$  1.4540**. It may contain up to 1% of the precursor 4-methoxybut-4-ene-2-one. It is easily purified by distilling through a Vigreux column in a vacuum and taking the middle fraction. [Danishefsky & Kitihara *J Am Chem Soc* **96** 7807 1974, Danishefsky *Acc Chem Res* **14** 400 1981.]

**Methylarsonic acid** [124-58-3] **M 139.9, m 159.8°, 161°,  $\text{pK}_1^{25}$  4.58,  $\text{pK}_2^{25}$  7.82** [ $\text{As}(\text{OH})_2$ ]. Recrystallise this herbicide from  $\text{Me}_2\text{CO}$  or absolute EtOH (white plates; its solubility is 28g/100ml EtOH). [Quick & Adams *J Am Chem Soc* **44** 809 1922, Beilstein **4** H 613, **4** I 577, **4** II 996, **4** III 1822, **4** IV 3682.]

**Methyl dichlorosilane (dichloro methylsilane)** [75-54-7] **M 115.0, m -92.5°, b 41°/748mm, 40.9°/760mm, 40-45°/atm,  $d_4^{27}$  1.105**. Impurities are generally other chloromethyl silanes. Distil it through a conventional Stedman column of 20 theoretical plates or more. It should be protected from  $\text{H}_2\text{O}$  by storing over  $\text{P}_2\text{O}_5$ . [Stock & Somieski *Chem Ber* **52** 695 1919, Sauer *J Am Chem Soc* **68** 962 1946, Beilstein **4** IV 4096.]

**Methyl Orange (sodium 4,4'-dimethylaminophenylazobenzenesulfonate)** [547-58-0] **M 327.3,  $\text{pK}_1^{25}$  3.56,  $\text{pK}_2^{25}$  6.49**. Recrystallise it twice from hot water, then wash it with a little EtOH followed by diethyl ether. It is an indicator: pH 3.1 (red) and pH 4.4 (yellow). [Beilstein **16** IV 510.]

**Methylphenyl dichlorosilane (dichloro methyl phenylsilane)** [149-74-6] **M 191.1, b 114-115°/50mm, 202-205°/atm,  $d_4^{20}$  1.17**. Purify it by fractionation using an efficient column. It hydrolyses *ca* ten times more slowly than methyltrichlorosilane and *ca* sixty times more slowly than phenyltrichlorosilane. [Shaffer & Flanigen *J Phys Chem* **61** 1591 1957, Beilstein **16** III 1211, **16** IV 1517.]

**Methylphosphonic acid** [993-13-5] **M 96.0, m 104-106°, 105-107°, 108°,  $\text{pK}_1^{25}$  2.12,  $\text{pK}_2^{25}$  7.29**. If it tests for  $\text{Cl}^-$ , then add  $\text{H}_2\text{O}$  and evaporate to dryness, repeat several times till free from  $\text{Cl}^-$ . The residue solidifies to a wax-like solid. *Alternatively*, dissolve the acid in the minimum volume of  $\text{H}_2\text{O}$ , add charcoal, warm, filter and evaporate to dryness in a vacuum over  $\text{P}_2\text{O}_5$ . [Kosolapoff *J Am Chem Soc* **75** 3379 1953.] The *di-Na salt* is prepared from 24g of acid in 50ml of dry EtOH, and a solution of 23g Na dissolved in 400ml EtOH is added. A white precipitate is formed, but the mixture is refluxed for 30 minutes to complete the reaction. Filter off the solid and recrystallise it from 50% EtOH. Dry the crystals in a vacuum desiccator. [Thompson *J Chem Soc* 3292 1952, Beilstein **4** IV 3498.]

**Methylphosphonic dichloride** [676-97-1] **M 132.9, m 33°, 33-37°, b 53-54°/10mm, 64-67°/20.5mm, 86°/44mm, 162°/760mm,  $d_4^{40}$  1.4382**. Fractionally redistil it until the purity as checked by hydrolysis and

acidimetry for Cl<sup>-</sup> is correct and the distillate should solidify on cooling. [Kinnear & Perren *J Chem Soc* 3437 1952, Crofts & Kosolapoff *J Am Chem Soc* 75 3379 1952, for IR see McIvor et al. *Can J Chem* 34 1611 1956, *Beilstein* 4 IV 3509.]

**Methyl Thymol Blue, sodium salt** {MTB, 3,3'-bis[*N,N'*-di(carboxymethyl)amino]thymolsulfonephthalein sodium salt} [1945-77-3] **M 844.8,  $\epsilon$  1.89 x 10<sup>4</sup> at 435nm, pH 5.5, pK<sub>1</sub><sup>25</sup> 3.0, pK<sub>2</sub><sup>25</sup> 3.3, pK<sub>3</sub><sup>25</sup> 3.8, pK<sub>4</sub><sup>25</sup> 7.4 (pK<sub>4</sub><sup>25</sup> 7.2), pK<sub>5</sub><sup>25</sup> 11.5 pK<sub>6</sub><sup>25</sup> 13.4.** The starting material for synthesis is Thymol Blue. Purify it as for Xylenol Orange in "Heterocyclic Compounds" in this Chapter. [Tereshin et al. *J Anal Chem USSR (Engl Trans)* 20 1138 1965, Körbl & Kakác *Col Czech Chem Commun* 23 889 1958, *Beilstein* 19/8 V 619.]

**Methyl trichlorosilane** [75-79-6] **M 149.5, b 13,7°/101mm, 64.3°/710.8mm, 65.5°/745mm, 66.1°/atm, d<sub>4</sub><sup>20</sup> 1.263, n<sub>D</sub><sup>20</sup> 1.4110.** If not very pure, distil it before use. The purity is checked by <sup>29</sup>Si NMR ( $\delta$  in MeCN is 13.14 ppm with respect to Me<sub>4</sub>Si). Possible contaminants are other silanes which can be removed by fractional distillation through a Stedman column of >72 theoretical plates with total reflux and 0.35% take-off. This apparatus should be under N<sub>2</sub> gas at a rate of 12 bubbles/minute fed into the line using an Hg manometer to control the pressure. It is sensitive to H<sub>2</sub>O. [Gillam et al. *J Am Chem Soc* 73 4252 1951, Olah et al. *J Org Chem* 48 3667 1983, *Beilstein* 4 IV 4212.]

**Methyl triethoxysilane** [2031-67-6] **M 178.31, b 142-144.5°/742mm, 141°/765mm, 141.5°/775mm, d<sub>4</sub><sup>20</sup> 0.8911, n<sub>D</sub><sup>20</sup> 1.3820.** Purify by fractionating it repeatedly in a stream of N<sub>2</sub> through a 3' Heligrid packed Todd column. It is hydrolysed by H<sub>2</sub>O and yields cyclic polysiloxanes on hydrolysis in the presence of acid in \*C<sub>6</sub>H<sub>6</sub>. [Bee et al. *J Am Chem Soc* 77 1292 1955, Sprung & Guenther *J Am Chem Soc* 77 3990 1955, *Beilstein* 4 IV 4204.]

**Methyl trimethoxysilane** [1185-55-3] **M 136.2, b 102°/760mm, d<sub>4</sub><sup>20</sup> 1.3687, n<sub>D</sub><sup>20</sup> 1.3711.** Likely impurities are 1,3-dimethyltetramethoxy disiloxane (b 31°/1mm) and cyclic polysiloxanes, see previous entry methyl triethoxysilane. [Tamborski & Post *J Org Chem* 16 1400 1952, Seyferth & Rochow *J Org Chem* 20 250 1955, *Beilstein* 4 IV 4203.]

***N*-Methyl-*N*-trimethylsilylacetamide** [7449-74-3] **M 145.3, b 48-49°/11mm, 84°/13mm, 105-107°/35mm (solid at room temperature), d<sub>4</sub><sup>20</sup> 0.90, n<sub>D</sub><sup>20</sup> 1.4379.** A likely impurity is Et<sub>3</sub>N.HCl which can be detected by its odour. If it is completely soluble in \*C<sub>6</sub>H<sub>6</sub>, then redistil; otherwise dissolve it in this solvent, filter and evaporate first in a vacuum at 12mm, then fractionate; all operations should be carried out in a dry N<sub>2</sub> atmosphere. [Klebe et al. *J Am Chem Soc* 88 3390 1966, Ried & Suarez-Rivero *Chem Ber* 96 1473 1963, *Beilstein* 4 IV 4011.]

**Methyl trimethylsilylacetate** [2916-76-9] **M 146.3, b 65-68°/50mm, d<sub>4</sub><sup>20</sup> 0.89.** Dissolve it in Et<sub>2</sub>O, shake with 1M HCl, wash with H<sub>2</sub>O, aqueous saturated NaHCO<sub>3</sub>, H<sub>2</sub>O again, and dry it (a precipitate may be formed in the NaHCO<sub>3</sub> solution and should be drawn off and discarded). The solvent is distilled off, and the residue is fractionated through a good column. The IR (CHCl<sub>3</sub>) has  $\nu_{\max}$  □□1728cm<sup>-1</sup>. [Fessenden & Fessenden *J Org Chem* 32 3535 1967, Matsuda et al. *J Org Chem* 45 237 1980, *Beilstein* 4 III 1855, 4 IV 3974 for Et ester.]

**Methyl 2-(trimethylsilyl)propionate** [55453-09-3] **M 160.3, b 155-157°/atm, d<sub>4</sub><sup>20</sup> 0.89.** Dissolve it in Et<sub>2</sub>O, wash it with aqueous NaHCO<sub>3</sub>, H<sub>2</sub>O, 0.1M HCl, H<sub>2</sub>O again, dry (MgSO<sub>4</sub>), evaporate and distil it. [Emde & Simchen *Synthesis* 867 1977, Oppolzer et al. *Tetrahedron* 39 3695 1983, Crimmin et al. *J Chem Soc Perkin Trans 1* 541 1985.]

***N*-Methyl-*N*-trimethylsilyl trifluoroacetamide** [24589-78-4] **M 199.3, b 78-79°/130mm.** Fractionate the amide through a 40mm Vigreux column. Usually it contains ca 1% of methyl trifluoroacetamide and 1% of other impurities which can be removed by gas chromatography or fractionating using a spinning band column. [Donike *J Chromatogr* 42 103 1969, Donike *J Chromatogr* 103 91 1975, *Beilstein* 4 IV 4011.]

**Methyl triphenoxyphosphonium iodide** [17579-99-6] **M 452.2, m 114° (sealed tube), 146°.** Gently heat the impure iodide with good grade Me<sub>2</sub>CO. The saturated solution obtained is decanted rapidly from undissolved

salt and treated with an equal volume of dry Et<sub>2</sub>O. The iodide separates as flat needles which are collected by centrifugation, washed several times with dry Et<sub>2</sub>O, and dried in a vacuum over P<sub>2</sub>O<sub>5</sub>. For this recrystallisation it is essential to minimise the time of contact with Me<sub>2</sub>CO and to work rapidly and with rigorous exclusion of moisture. If the crude material is to be used, it should be stored under dry Et<sub>2</sub>O, and dried and weighed *in vacuo* immediately before use. [Landauer & Rydon *J Chem Soc* 224 1953, Hudson et al. *J Chem Soc Perkin Trans I* 982 1974, *Beilstein* 6 IV 704.]

**Methyl triphenylphosphonium bromide** [1779-49-3] **M 357.3, m 229-230°(corr), 227-229°, 230-233°**. If the solid is sticky, wash it with \*C<sub>6</sub>H<sub>6</sub> and dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. [Marvel & Gall *J Org Chem* 24 1494 1959, *Chem Ber* 87 1318 1954, Milas & Priesing *J Am Chem Soc* 79 6295 1957, Wittig & Schöllkopf *Org Synth* 40 66 1960.] The *iodide*, on recrystallisation from H<sub>2</sub>O, has **m 187.5-188.5°** [Mann et al. *J Chem Soc* 1130 1953, Wittig & Geissler *Justus Liebigs Ann Chem* 580 44 1953]. [*Beilstein* 16 IV 981.]

**Methyl vinyl dichlorosilane (dichloro methyl vinyl silane)** [124-70-9] **M 141.1, b 43-45.5°/11-11.5mm, 91°/742mm, 92.5°/743.2mm, 92.5-93°/atm, d<sub>4</sub><sup>20</sup> 1.0917, n<sub>D</sub><sup>20</sup> 1.444**. Likely impurities are dichloromethylsilane, butadienyl-dichloromethylsilane. Fractionate the silane through a column packed with metal filings (20 theoretical plates) at atmospheric or reduced pressure. [*Izv Akad Nauk SSSR Ser Khim* 1474 1957 and 767 1958, *Beilstein* 4 III 1894, 4 IV 4184.]

**Milling Red SWB {1-[4-[4-[4-toluenesulfonyloxy]phenylazo](3,3'-dimethyl-1,1'-biphenyl)-4'-azo]-2-hydroxynaphthalene-6,8-disulfonic acid di-Na salt, Acid Red 114}** [6459-94-5] **M 830.8, m dec >250°, CI 23635, λ<sub>max</sub> ~514nm**. Salt out three times with sodium acetate, then repeatedly extract it with EtOH and dry the solid in air. [McGrew & Schneider *J Am Chem Soc* 72 2547 1950, *Beilstein* 16 II 140.] See Solochrome Violet R [2092-55-9] in "Aromatic compounds" in this Chapter.

**Milling Yellow G** [51569-18-7]. Salted out three times with sodium acetate, then repeatedly extracted with EtOH. [McGrew & Schneider *J Am Chem Soc* 72 2547 1950, *Beilstein* 16 II 125.] See Solochrome Violet R [2092-55-9] in "Aromatic compounds" in this Chapter.

**Naphthalene Scarlet Red 4R [1-(4-sulfonaphthalene-1-azo)-2-hydroxynaphthalene-6,8-disulfonic acid tri-Na salt, New Coccine, Acid Red 18]** [2611-82-7] **M 604.5, m >250°(dec), CI 16255, λ<sub>max</sub> 506nm**. Dissolve the dye in the minimum quantity of boiling water, filter and enough EtOH is added to precipitate *ca* 80% of the dye. This process is repeated until a solution of the dye in aqueous 20% pyridine has a constant extinction coefficient. [*Beilstein* 16 I 306.]

**Naphthol Yellow S (citronin A, flavianic acid sodium salt, 8-hydroxy-5,7-dinitro-2-naphthalene sulfonic acid disodium salt)** [846-70-8] **M 358.2, decomposes on heating**. It is a water-soluble greenish yellow powder. The *free sulfonic acid* can be recrystallised from dilute HCl (m 150°) or AcOH/EtOAc (m 148-149.5°). The disodium salt is then obtained by dissolving the acid in two equivalents of aqueous NaOH and evaporating to dryness and drying the residue in a vacuum desiccator. The sodium salt can be recrystallised from the minimum volume of H<sub>2</sub>O or from EtOH [Dermer & Dermer *J Am Chem Soc* 61 3302 1939]. [*Beilstein* 11 III 542.]

**1,2-Naphthoquinone-4-sulfonic acid sodium salt (3,4-dihydro-3,4-dioxo-1-naphthlene sulfonic acid sodium salt)** [521-24-4] **M 260.2, pK<sub>Est</sub> <0**. It forms yellow crystals from aqueous EtOH and should be dried at 80° *in vacuo*. Its solubility in H<sub>2</sub>O is 5% [Martin & Fieser *Org Synth Coll Vol III* 633 1955, Danielson *J Biol Chem* 101 507 1933, UV: Rosenblatt et al. *Anal Chem* 27 1290 1955]. [*Beilstein* 11 IV 668.]

**1-Naphthyl phosphate disodium salt** [2183-17-7] **M 268.1, pK<sub>1</sub><sup>25</sup> 0.97, pK<sub>2</sub><sup>25</sup> 5.85 (for free acid)**. The free acid has **m 157-158°** (from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>). The *free acid* is recrystallised several times by adding 20 parts of boiling \*C<sub>6</sub>H<sub>6</sub> to a hot solution of 1 part of *free acid* and 1.2 parts of Me<sub>2</sub>CO. It has **m 157-158°**. [Chanley & Ferguson *J Am Chem Soc* 77 4002 1955.] The *monosodium salt* is precipitated from a solution of the acid phosphate in MeOH by addition of an equivalent of MeONa in MeOH. [Friedman & Seligman *J Am Chem Soc*

72 624 1950, *Beilstein* 6 IV 4226.] See also entry on p 877 in Chapter 7.

**2-Naphthyl phosphate monosodium salt** [14463-68-4] M 246.1, m 203-205°, pK<sub>1</sub><sup>25</sup> 1.25, pK<sub>2</sub><sup>25</sup> 5.83 (for free acid). Recrystallise it from H<sub>2</sub>O (10ml) containing NaCl (0.4g). The salt is collected by centrifugation and dried in a vacuum desiccator, m 203-205° (partially resolidifies and re-melts at 244°). It crystallises from MeOH (m 222-223°). The free acid is recrystallised several times by addition of 2.5 parts of hot CHCl<sub>3</sub> to a hot solution of the free acid (1 part) in Me<sub>2</sub>CO (1.3 parts), m 177-178°. [Friedman & Seligmann *J Am Chem Soc* 73 5292 1951, Chanley & Feaggeson *J Am Chem Soc* 77 4002 1955, *Beilstein* 6 H 647, 6 III 2989, 6 IV 428.] See also entry on p 877 in Chapter 7.

**2-Nitrophenol-4-arsonic acid (4-hydroxy-3-nitrophenylarsonic acid)** [121-19-7] M 263.0, pK<sub>Est(1)</sub>~4.4 As(O)-(OH)-(O<sup>-</sup>), pK<sub>Est(2)</sub>~7.4 (phenolic OH), pK<sub>Est(3)</sub>~7.7 (As(O)-2(O<sup>-</sup>)). It crystallises from water and is used for the spectroscopic detection of Zn. [*Beilstein* 16 II 468, 16 III 1073, 16 IV 1188.]

**1-Nitroso-2-naphthol-3,6-disulfonic acid, di-Na salt, hydrate (Nitroso-R-salt)** [525-05-3] M 377.3, m >300°, pK<sub>Est(1)</sub><0 (SO<sub>3</sub><sup>-</sup>), pK<sub>Est(2)</sub>~7 (OH). Purify the salt by dissolution in aqueous alkali and precipitation by addition of HCl. The pKa is 7.13. [Oka & Miyamoto *J Chem Soc Jpn* (Pure Chem Sectn) 76 672 1955 (*Chem Abstr* 11882 1956I), *Beilstein* 11 IV 669.]

**Nuclear Fast Red (1-amino-2,4-dihydroxy-5,10-anthraquinone-3-sulfonic acid Na Salt)** [6409-77-4] M 357.3, m >290°(dec), λ<sub>max</sub> 518nm. A solution of 5g of the dye in 250ml of warm 50% EtOH is cooled to 15° for 36 hours, then filtered on a Büchner funnel, washed with EtOH until the washings are colourless, then with 100ml of Et<sub>2</sub>O and dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. It is a biological stain that is also used for the estimation of Ca. [Kingsley & Robnett *Anal Chem* 33 552 1961.]

**Octadecyl trichlorosilane** [112-04-9] M 387.9, b 159-162°/13mm, 185-199°/2-3mm, d<sub>4</sub><sup>30</sup> 0.98. Purify it by fractional distillation at high vacuum. [Winstein & Seubold *J Am Chem Soc* 69 2916 1947, *Beilstein* 4 IV 4256.]

**Octamethyl cyclotetrasiloxane** [556-67-2] M 296.6, m 17-19°, 17.58°, 18.5°, b 74°/20mm, 176.4°/760mm, d<sub>4</sub><sup>29.3</sup> 0.9451, n<sub>D</sub><sup>30</sup> 1.3968. The solid exists in two forms, m 16.30° and 17.65°. Dry it over CaH<sub>2</sub> and distil it. Further fractionation can be effected by repeated partial freezing and discarding the liquid phase. [Osthoff & Grubb *J Am Chem Soc* 76 399 1954, Hoffman *J Am Chem Soc* 75 6313 1953, *Beilstein* 4 IV 4125.]

**Octamethyl trisiloxane** [107-51-7] M 236.5, m -80°, b 151.7°/747mm, 153°/760mm. Distil it twice, the middle fraction from the first distillation is again distilled, and the middle fraction of the second distillation is used. [Patnode & Wilcock *J Am Chem Soc* 68 358, Wolcock *J Am Chem Soc* 68 691 1946, Thompson *J Chem Soc* 1908 1953, *Beilstein* 4 IV 4115.]

**Octaphenyl cyclotetrasiloxane** [546-56-5] M 793.2, m 201-202°, 203-204°, b 330-340°/1mm. Recrystallise it from AcOH, EtOAc, \*C<sub>6</sub>H<sub>6</sub> or \*C<sub>6</sub>H<sub>6</sub>/EtOH. It forms two stable isomorphs and both forms, as well as the mixture, melt at 200-201°. There is a metastable form which melts at 187-189°. [Burkhard et al. *J Am Chem Soc* 67 2174 1945, Hyde et al. *J Am Chem Soc* 69 488 1947, *Beilstein* 16 IV 1530.]

**Octyl trichlorosilane** [5283-66-9] M 247.7, b 96.5°/10mm, 112°/15mm, 119°/28mm, 229°/760mm, d<sub>4</sub><sup>20</sup> 1.0744, n<sub>D</sub><sup>20</sup> 1.4453. Purify the silane by repeated fractionation using a 15-20 theoretical plate glass column packed with glass helices. This can be done more efficiently using a spinning band column. Its purity can be checked by analysing for HCl (*ca* 0.5-1g of sample is dissolved in 25ml of MeOH, diluted with H<sub>2</sub>O and the HCl formed by hydrolysis is titrated with standard alkali). It is moisture sensitive. [Whitmore *J Am Chem Soc* 68 475 1946, El-Abbadly & Anderson *J Am Chem Soc* 80 1737 1958, *Beilstein* 4 III 1907.]

**Orange I [tropaeolin 000 Nr1, 4-(4-hydroxy-1-naphthylazo)benzenesulfonic acid sodium salt]** [523-44-4] M 350.3, m >260°(dec). Purify the dye by dissolving it in the minimum volume of H<sub>2</sub>O, adding, with stirring, a

large excess of EtOH. The salt separates as orange needles. It is collected by centrifugation or filtration, washed with absolute EtOH (3x) and Et<sub>2</sub>O (2x) in the same way, and dried in a vacuum desiccator over KOH. The *free acid* can be recrystallised from EtOH. [Slotta & Franke *Chem Ber* **64** 86 1931, *Beilstein* **16** H 275, **16** II 117, **16** IV 410.] The purity can be checked by titration with titanium chloride [Klotz *J Am Chem Soc* **68** 2299 1946].

**Orange II [tropaeolin 000 Nr2, 4-(2-hydroxy-1-naphthylazo)benzenesulfonic acid sodium salt]** [633-96-5] **M 350.3**. Purification is as for Orange I. Its solubility in H<sub>2</sub>O is 40g/L at 25°. [Müller et al. *Helv Chim Acta* **35** 2579 1952.] Also purify it by extracting it with a small volume of cold water, then crystallising it by dissolving in boiling water, cooling to *ca* 80°, adding two volumes of EtOH and cooling. When cold, the precipitate is filtered off, washed with a little EtOH and dried in air. It can be salted out from aqueous solution with sodium acetate, then repeatedly extracted with EtOH. Meggy and Sims [*J Chem Soc* 2940 1956], after crystallising the sodium salt twice from water, dissolved it in cold water (11ml/g) and added conc HCl to precipitate the *acid dye* which was separated by centrifugation, redissolved and again precipitated with acid. After washing the precipitate three times with 0.5M acid, it was dried over NaOH, recrystallised twice from absolute EtOH, washed with a little Et<sub>2</sub>O, dried over NaOH and stored over conc H<sub>2</sub>SO<sub>4</sub> in the dark. It can then be converted to the pure salt with the calculated amount of NaOH or Na<sub>2</sub>CO<sub>3</sub>. [*Beilstein* **16** IV 408.]

**Orange G (1-phenylazo-2-naphthol-6,8-disulfonic acid di-Na salt)** [1936-15-8] **M 452.4**, **pK<sub>Est</sub>~9**. Recrystallise this dye from 75% EtOH, dry it for 3 hours at 110° and keep it in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. The *free acid* crystallises from EtOH or conc HCl in deep red needles with a green reflex. [Conant & Pratt *J Am Chem Soc* **48** 2483 1923, Drew & Landquist *J Chem Soc* 292 1938, *Beilstein* **16** H 301, **16** I 305, **16** II 141, **16** III 327.]

**Orange RO {acid orange 8, 1,8-[bis(4-*n*-propyl-3-sulfophenyl-1-amino)]anthra-9,10-quinone di-Na salt}** [5850-86-2] **M 364.4**, **CI 15575**, **λ<sub>max</sub> 490nm**. Salt it out three times with sodium acetate, then extract it repeatedly with EtOH and dry *in vacuo*.

**Pentafluorophenyl dimethylchlorosilane (Flopemesyl chloride)** [20082-71-7] **M 260.7**, **b 89-90°/10mm**, **d<sub>4</sub><sup>30</sup> 1.403**, **n<sub>D</sub><sup>30</sup> 1.447**. If it goes turbid on cooling due to separation of some LiCl, then dissolve it in Et<sub>2</sub>O, filter and fractionate it in a vacuum. [Morgan & Poole *J Chromatogr* **89** 225 1974, Birkinshaw et al. *J Chromatogr* **132** 548 1977.]

**Phenylarsonic acid (benzenearsonic acid)** [98-05-5] **M 202.2**, **m 155-158°(dec)**, **158-162°(dec)**, **pK<sub>1</sub><sup>25</sup> 3.65**, **pK<sub>2</sub><sup>25</sup> 8.77**. Crystallise it from H<sub>2</sub>O (3ml/g) between 90° and 0°. *Alternatively*, dissolve 600g of the acid in 500ml of boiling H<sub>2</sub>O, add 20g of Norite, filter hot, cool, filter off the crystals and dry them. On heating at ~154-160° it is converted to the *anhydride*. [Bullard & Dickery *Org Synth Coll Vol II* 494 1943, for the 4-nitro derivative see Ruddy & Starkey *Org Synth Coll Vol III* 665 1955, *Beilstein* **16** H 868, **16** I 448, **16** II 457, **16** III 1057, **16** IV 1183.]

**Phenylboronic acid (benzeneboronic acid)** [98-80-6] **M 121.9**, **m ~43°, 215-216° (anhydride)**, **217-220°**, **pK<sub>1</sub><sup>25</sup> 8.83**. It recrystallises from H<sub>2</sub>O, but it can convert spontaneously to *benzeneboronic anhydride* or phenylboroxide on standing in dry air. A possible impurity is dibenzeneborinic acid which can be removed by washing with petroleum ether. Heating in an oven at 110°/760mm for 1 hour converts it to the *anhydride* **m** 214-216°. Its solubility in H<sub>2</sub>O is 1.1% at 0° and 2.5% at 25°, and in EtOH it is 10% (w/v). [Gilman & Moore *J Am Chem Soc* **80** 3609 1958.] If the acid is required, not the anhydride, the acid (from recrystallisation in H<sub>2</sub>O) is dried in a slow stream of air saturated with H<sub>2</sub>O. The *anhydride* is converted to the acid by recrystallisation from H<sub>2</sub>O. The acid gradually dehydrates to the *anhydride* if left in air at room temperature with 30-40% relative humidity. The melting point is usually that of the anhydride because the acid dehydrates before it melts [Washburn et al. *Org Synth Coll Vol IV* 68 1963]. [*Beilstein* **16** IV 1654.]

**1,2-Phenylenephosphorochloridate (2-chloro-1,3,2-benzodioxaphosphole-2-oxide)** [1499-17-8] **M 190.5**, **m 52°, 58-59°, 59-61°**, **b 80-81°/1-2mm**, **118°/10mm**, **122°/12mm**, **125°/16mm**, **155°/33mm**. After distilling it in a vacuum, it sets to a colourless solid. It is soluble in petroleum ether, \*benzene and slightly soluble in Et<sub>2</sub>O.

[Khwaja et al. *J Chem Soc (C)* 2092 1970, Anschütz & Broeker *Justus Liebigs Ann Chem* **454** 109 1927, *Beilstein* **6** IV 5602.]

**Phenylphosphinic acid [benzenephosphinic acid, PhPH(O)(OH)]** [1779-48-2] **M 142.1, m 70°, 71°, 83-85°, 86°, pK<sub>1</sub><sup>25</sup> 1.75.** Crystallise it from H<sub>2</sub>O (solubility is 7.7% at 25°). Also purify it by placing the solid in a flask covered with dry Et<sub>2</sub>O, and allowed it to stand for 1 day with intermittent shaking. Et<sub>2</sub>O is decanted off and the process repeated. After filtration, excess Et<sub>2</sub>O is removed in a vacuum. It has also been recrystallised from \*C<sub>6</sub>H<sub>6</sub>. [Michaelis *Justus Liebigs Ann Chem* **181** 265 1876, Banks & Skoog *Anal Chem* **29** 109 1957, NMR: Van Wazer et al. *J Am Chem Soc* **78** 5715 1956, *Beilstein* **16** IV 1033.]

**Phenylphosphonic acid** [1571-33-1] **M 158.1, m 164.5-166°, pK<sub>2</sub><sup>25</sup> 7.43 (7.07).** It is best to recrystallise it from H<sub>2</sub>O by concentrating an aqueous solution to a small volume and allowing it to crystallise. Wash the crystals with ice cold H<sub>2</sub>O and dry them in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub>. [Lecher et al. *J Am Chem Soc* **76** 1045 1954.] pK<sub>2</sub><sup>25</sup> values in H<sub>2</sub>O are 7.07, and in 50% EtOH 8.26. [Jaffé et al. *J Am Chem Soc* **75** 2209 1953, IR: Daasch & Smith *Anal Chem* **23** 853 1951, *Beilstein* **16** IV 1068.]

**Phenylphosphonic dichloride (P,P-dichlorophenyl phosphine oxide)** [824-72-6] **M 195.0, m 3°, b 83-84°/1mm, 135-136°/23mm, d<sub>4</sub><sup>30</sup> 1.977, n<sub>D</sub><sup>30</sup> 1.5578.** Fractionally distil it using a very efficient or a spinning band column. [Lecher et al. *J Am Chem Soc* **76** 1045 1954, NMR: Müller et al. *J Am Chem Soc* **78** 3557 1956, Van Wazer et al. *J Am Chem Soc* **78** 5715 1956, IR: Daasch & Smith *Anal Chem* **23** 853 1951, *Beilstein* **16** IV 1074.]

**Phenylphosphonous acid [PhP(OH)<sub>2</sub>] [tautomer of phenylphosphinic acid – above]** [121-70-0] **M 141.1, m 71°, pK<sub>Est</sub> <0, pK<sub>17</sub> 2.1.** Crystallise it from hot H<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>. [*Beilstein* **16** IV 1033.]

**Phenylphosphonous dichloride (P,P-dichloro phenyl phosphine)** [644-97-3] **M 179.0, m -51°, b 68-70°/1mm, 224-226°/atm, d<sub>4</sub><sup>30</sup> 1.9317, n<sub>D</sub><sup>35</sup> 1.5962.** Distil it in a vacuum by fractionating through a 20cm column packed with glass helices (better to use a spinning band column) [Buchner & Lockhart *J Am Chem Soc* **73** 755 1951, NMR: Müller et al. *J Am Chem Soc* **78** 3557 1956, IR: Daasch & Smith *Anal Chem* **23** 853 1951]. It forms a yellow *Ni* complex: Ni(C<sub>6</sub>H<sub>5</sub>Cl<sub>2</sub>P)<sub>4</sub> (**m** 91-92°, from H<sub>2</sub>O) [Quin *J Am Chem Soc* **79** 3681 1957] and a yellow complex also with molybdenum carbonyl: Mo(CO)<sub>3</sub>.(C<sub>6</sub>H<sub>5</sub>Cl<sub>2</sub>P)<sub>3</sub> (**m** 106-110° dec) [Abel et al. *J Chem Soc* 2323 1959]. [*Beilstein* **16** IV 972.]

**Phenyl phosphoryl dichloride** [770-12-7] **M 211.0, m -1°, b 103-104°/2mm, 110-111°/10mm, 130-134°/21mm, 241-243°/atm, d<sub>4</sub><sup>30</sup> 1.4160, n<sub>D</sub><sup>30</sup> 1.5216.** Fractionally distil it under as high a vacuum as possible using an efficient fractionating column or a spinning band column. It should be redistilled if the IR is not very satisfactory [IR: Bellamy & Beecher *J Chem Soc* 475 1952, Freeman & Colver *J Am Chem Soc* **60** 750 1938, Orloff et al. *J Am Chem Soc* **80** 727 1958]. [*Beilstein* **6** IV 737.] **HARMFUL VAPOURS.**

**Phenylsilane (PhSiH<sub>3</sub>)** [694-53-1] **M 108.2, b 62°/100mm, 120°/atm, d<sub>4</sub><sup>25</sup> 0.877, n<sub>D</sub><sup>20</sup> 1.5125.** It is best prepared by reduction of PhSiCl<sub>3</sub> with LiAlH<sub>4</sub>/Et<sub>2</sub>O: stir at room temperature overnight, pour onto crushed ice, extract with Et<sub>2</sub>O, dry the extract (Drierite), filter, evaporate and distil the residual oil preferably in a vacuum [Beukesser et al. *J Am Chem Soc* **74** 648 1952, Koch et al. *Rec Trav Chim Pays Bas* **114** 206 1995, Finholt et al. *J Am Chem Soc* **69** 2692 1947].

It is a useful reagent for deoxygenating phosphine oxides to phosphines, and will reduced chiral phosphine oxides to chiral phosphines with retention of configuration [Marsi *J Org Chem* **39** 265 1974]. It is a reducing agent in radical reactions [Perez et al. *J Org Chem* **52** 5570 1967], and deoxygenates primary and secondary alcohols *via* their xanthate or thionocarbonate esters (prepared from aryl-OCSCl and an inhibitor, e.g benzoyl peroxide) in 60-100 minutes in >87% yields [Barton et al. *Synlett* 435 1991]. [*Beilstein* **16** III 1198, **16** IV 1360.]

**Phenylthio trimethylsilane (trimethyl phenylthio silane)** [4551-15-9] **M 182.4, b 95-99°/12mm, d<sub>4</sub><sup>30</sup> 0.97.** Purification is as for phenyl trimethylsilylmethyl sulfide.

**Phenyl trimethoxysilane (trimethoxysilyl benzene)** [2996-92-1] **M 198.3, b 103°/20mm, 130.5-131°/45mm,  $d_4^{35}$  1.022,  $n_D^{35}$  1.4698.** Fractionate it through an efficient column but note that it forms an azeotrope with MeOH which is a likely impurity. [Kantor *J Am Chem Soc* **75** 2712 1953 *Beilstein* **16** IV 1556.]

**Phenyl trimethylsilylmethyl sulfide [(phenylthiomethyl)trimethylsilane]** [17873-08-4] **M 196.4, b 48°/0.04mm, 113-115°/12mm, 158.5°/52mm,  $d_4^{30}$  0.9671,  $n_D^{30}$  1.5380.** If the sample is suspect, then add H<sub>2</sub>O, wash it with 10% aqueous NaOH, H<sub>2</sub>O again, dry (anhydrous CaCl<sub>2</sub>) and fractionally distil it through a 2ft column packed with glass helices. [Cooper *J Am Chem Soc* **76** 3713 1954.]

**Phosphine** [7803-51-2] **M 34.0, m -133°, b -87.5°/760mm, -87.7°/760mm, critical temperature 51.3°,  $d_4^{25}$  1.5307,  $pK^{25}$  ~ -14 (extremely weak base).** Phosphine is a gas with a very strong odour of fish and is **POISONOUS**. The gas is poorly soluble in H<sub>2</sub>O (0.26ml/ml at 20°), and ignites spontaneously in air with a luminous glow. It has been prepared in various ways. A convenient preparation is to make aluminium phosphide by mixing 2 parts of Al powder and 1 part of red P on a piece of paper. By igniting the paper, the mixture becomes white-hot resulting in a spongy mass of aluminium phosphide [Hoffman *J Am Chem Soc* **43** 1684 1921, Bodoux *Bull Soc Chim Fr* **27** (3) 568 1902]. This phosphide reacts with cold H<sub>2</sub>O to give a steady stream of PH<sub>3</sub>. [Use an efficient fume cupboard in these experiments.] Fortunately, the gas is available in metal cylinders, but all due precautions should be taken. [Klement in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 525-530 1963, Gokhale and Jolly *Inorg Synth* **9** 56 1967, for basicity see Henderson & Streuli *J Am Chem Soc* **82** 5791 1960.]

**Polystyrenesulfonic acid sodium salt (-CH<sub>2</sub>CH(C<sub>6</sub>H<sub>4</sub>SO<sub>3</sub>Na)-)** [25704-18-1]. Purify the polymer by repeated precipitation of the sodium salt from an aqueous solution by MeOH, with subsequent conversion to the free acid by passage through an Amberlite IR-120 ion-exchange resin. [Kotin & Nagasawa *J Am Chem Soc* **83** 1026 1961.] Recrystallise it from EtOH. *Alternatively*, purify it by passage through cation and anion exchange resins in series (Rexyn 101 cation exchange resin and Rexyn 203 anion exchange resin), then titrated it with NaOH to pH 7. The sodium form of polystyrenesulfonic acid is precipitated by addition of 2-propanol. Dry it in a vacuum oven at 80° for 24 hours, and finally increasing to 120° before to use. [Kowblansky & Ander *J Phys Chem* **80** 297 1976.]

**Pontacyl Carmine 2G (Acid Red 1, Amido Naphthol Red G, Azophloxine, 1-acetamido-8-hydroxy-7-phenylazonaphthalene-3,7-disulfonic acid di-Na salt)** [3734-67-6] **M 510.4, CI 18050,  $\lambda_{max}$  532nm.** Salt it out three times with sodium acetate, then repeatedly extract it with EtOH. See Solochrome Violet R [2092-55-9] in "Aromatic compounds" in this Chapter. [McGrew & Schneider *J Am Chem Soc* **72** 2547 1950.]

**Pontacyl Light Yellow GX [Acid Yellow 17, 1-(2,5-dichloro-4-sulfophenyl)-3-methyl-4-(4-sulfophenylazo)-5-hydroxypyrazole di-Na Salt]** [6359-98-4] **M 551.3, CI 18965,  $\lambda_{max}$  400nm.** Purify as for Pontacyl Carmine 2G above.

**Potassium 4-acetylphenyltrifluoroborate** [252726-24-2] **M 226.1, m >250°, 290°(dec).** The salt is prepared by adding an aqueous solution of KHF<sub>2</sub> (41ml, 4.5M solution, 185mmol) to a solution of 4-acetylphenylboronic acid (10g, 61mmol) in MeOH (40ml) at ~25°, when a heavy precipitate deposits, but the suspension is stirred for 1 hour at 25° and the solid is filtered off, washed with MeOH and recrystallised from the minimum volume of Me<sub>2</sub>CO to provide *p*-AcC<sub>6</sub>H<sub>4</sub>-BF<sub>3</sub><sup>-</sup> K<sup>+</sup> (12, 87%), m >250°. Its <sup>1</sup>H NMR [(CD<sub>3</sub>)<sub>3</sub>CO, 500MHz] has  $\delta$  at 7.74(d, *J* = 7.8Hz, 2H), 7.59 (d, *J* = 7.8Hz, 2H), 2.49 (s, 3H); the <sup>13</sup>C NMR [(CD<sub>3</sub>)<sub>3</sub>CO, 125MHz] has  $\delta$  at 198.1, 134.2, 131.3 (d, *J* = 1.5Hz), 126.2, 26.4; the <sup>19</sup>F NMR (DMSO-*d*<sub>6</sub>, 470MHz) has  $\delta$  at -140.3 (br s), and the <sup>11</sup>B NMR (DMSO-*d*<sub>6</sub>, 64MHz) has  $\delta$  at 3.27 (br s), and has the correct elemental analysis for C and H. It has been used for inserting a 4-acetylphenyl group into the position of a halogen or triflate group of various arenes and various heterocycles in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *J Org Chem* **68** 4302 2003.]

**Potassium 3,5-bis(trifluoromethyl)phenyltrifluoroborate** [166328-09-2] **M 320.0, m >250°.** This salt was prepared by the same procedure as the preceding salt in 89% yield and recrystallised from the minimum volume of Me<sub>2</sub>CO. Its <sup>1</sup>H NMR (DMSO-*d*<sub>6</sub>, 500MHz) has  $\delta$  at 7.87(s, 2H), 7.71 (s, 1H), 2.49 (s, 3H); the <sup>13</sup>C NMR

(DMSO- $d_6$ , 125MHz) has  $\delta$  at 131.2, 128.1 (q,  $J = 31\text{Hz}$ ), 124.2 (q,  $J = 272\text{Hz}$ ), 118.6 (s); the  $^{19}\text{F}$  NMR (DMSO- $d_6$ , 470MHz) has  $\delta$  at -61.7, -141.6 (br d,  $J = 71\text{Hz}$ ), and the  $^{11}\text{B}$  NMR (DMSO- $d_6$ , 64MHz) has  $\delta$  at 2.57 (br s), and has the correct elemental analysis for C and H. It has been used for inserting a 3,5-bis(trifluoromethyl)phenyl group into the position of a halogen or triflate group of various arenes and various heterocycles in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *J Org Chem* **68** 4302 2003, Vedejs et al. *J Org Chem* **60** 3020 1995.]

**Potassium 2,6-difluorophenyltrifluoroborate** [267006-25-7] **M 220.0, m <300°**. This salt was prepared by the same procedure as the preceding bis(trifluoromethyl) salt in 90% yield except that the crystalline precipitate was washed with  $\text{H}_2\text{O}$  then  $\text{Et}_2\text{O}$  and dried at high vacuum in a Schlenk line, and had  $m >250^\circ$ . Its  $^1\text{H}$  NMR (DMSO- $d_6$ , 500MHz) has  $\delta$  at 6.66 (m, 2H), 7.10 (m, 1H); the  $^{13}\text{C}$  NMR (DMSO- $d_6$ , 125MHz) has  $\delta$  at 168.7 (dd,  $J = 242\text{Hz}$ , 18H), 127.5 (t,  $J = 11\text{Hz}$ ), 110.0 (dd,  $J = 23\text{Hz}$ , 8H), 118.6 (s); the  $^{19}\text{F}$  NMR (DMSO- $d_6$ , 470MHz) has  $\delta$  at -103.7 (q,  $J = 9.4\text{Hz}$ ), -132.6 (qt,  $J = 43.9\text{Hz}$ ); and the  $^{11}\text{B}$  NMR (DMSO- $d_6$ , 64MHz) has  $\delta$  at 2.17 (q,  $J = 44\text{Hz}$ ), and has the correct elemental analysis for C and H. It has been used for inserting a 2,6-difluorophenyl group into the position of a halogen or triflate group of various arenes and various heterocycles in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *J Org Chem* **68** 4302 2003.]

**Potassium 2,4-difluorophenyltrifluoroborate** [871231-41-3] **M 220.0, m >300°** can be prepared by the same procedure as above and applied in similar reactions.

**Potassium 2,6-dimethylphenyltrifluoroborate** [561328-67-4] **M 212.1, m >250°**. This salt was prepared by the same procedure as the preceding difluorophenyl salt in 88% yield except that the crystalline precipitate was washed with  $\text{H}_2\text{O}$  then  $\text{Et}_2\text{O}$  and dried at high vacuum in a Schlenk line and had  $m >250^\circ$ . Its  $^1\text{H}$  NMR (acetone- $d_6$ , 500MHz) has  $\delta$  at 6.83 (t,  $J = 7.4\text{Hz}$ , 1H), 6.75 (d,  $J = 7.4\text{Hz}$ , 2H), 2.40 (s, 6H); the  $^{13}\text{C}$  NMR (acetone- $d_6$ , 125MHz) has  $\delta$  at 142.5, 127.7, 125.9, 23.8; the  $^{19}\text{F}$  NMR (acetone- $d_6$ , 470MHz) has  $\delta$  at -132.5 (q,  $J = 48\text{Hz}$ ); and the  $^{11}\text{B}$  NMR (acetone- $d_6$ , 64MHz) has  $\delta$  at 4.94 (br d,  $J = 49\text{Hz}$ ), and has the correct elemental analysis for C and H. It has been used for inserting a 2,6-dimethylphenyl group into the position of a halogen or triflate group of various arenes and various heterocycles in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *J Org Chem* **68** 4302 2003, cf Vedejs et al. *J Org Chem* **60** 3020 1995.] **Potassium 4-tert-butylphenyltrifluoroborate** [423118-47-2] **M 240.1, m >230°**, can be similarly prepared from 4-tert-butylphenylboronic acid and used in Pd catalysed Suzuki cross-coupling reactions.

**Potassium ethyl xanthate (potassium O-ethyl dithiocarbonate, potassium ethyl xanthogenate)** [140-89-6] **M 160.3, m ~210°(dec), > 215°(dec), pK<sup>25</sup> 2.16 (for -S<sup>-</sup>)**. It forms white to pale yellow crystals, prepared by adding ethanolic KOH dropwise to an ethanolic solution of  $\text{CS}_2$  to form the salt which can be coaxed further out of solution by adding  $\text{Et}_2\text{O}$ . Crystallise it from absolute EtOH, ligroin/ethanol or acetone by adding  $\text{Et}_2\text{O}$ . Wash it with ether, then dry it in a desiccator. Its solubility at  $56^\circ$  in EtOH is  $>1\%$  and in  $\text{Me}_2\text{CO}$  is 8% (the Na salt has 44%), and very high in  $\text{H}_2\text{O}$  at  $25^\circ$ . Dry it *in vacuo*, if it contains  $\text{H}_2\text{O}$ , and store it in a tightly stoppered bottle away from light. [Warren & Matthews *Anal Chem* **26** 1985 1954, *Beilstein* **3** H 209, **3** I 84 **3** II 152, **3** III 336, **3** IV 402.]

**Potassium 2-furantrifluoroborate** [166328-14-9] **M 174.0, m 200°(dec), >293-303°**. This salt has been prepared from furan (5.0ml, 68.7mmol, dried over 3Å Molecular Sieves) in dry THF (50ml) which was treated under  $\text{N}_2$  with BuLi (42.0ml, 1.64M in pentane, 68.9mmol) and stirred at  $-5^\circ$  for 3.5 hours to form furyllithium. The latter was treated with  $\text{B}(\text{iso-PrO})_3$  and allowed to warm to  $\sim 25^\circ$ , quenched with 10% aqueous HCl ( $\sim 50\text{ml}$ ), diluted with  $\text{Et}_2\text{O}$  (50ml) and the organic layer was extracted with 1N NaOH (2 X 50ml). The alkaline extracts were combined, acidified to pH 3 with 10% aqueous HCl, and the acid layer was then extracted with  $\text{Et}_2\text{O}$  (3 X 50ml). The combined  $\text{Et}_2\text{O}$  extracts were dried ( $\text{Na}_2\text{SO}_4$ ) and evaporated to dryness. The residual 2-furanboronic acid was dissolved in MeOH (200ml) and  $\text{H}_2\text{O}$  (40ml) was treated with 3 equivalents of  $\text{KHF}_2$  (16g, 206mmol), refluxed overnight and evaporated to dryness *in vacuo*. The residue was extracted with MeCN (2 X 30ml), filtered and the filtrate was evaporated to dryness, the residue was washed with  $\text{Et}_2\text{O}$ , dried and recrystallised from MeCN/EtOAc to give pure 2-furan- $\text{BF}_3^- \text{K}^+$  (5.8g, 48%) as yellow crystals. Attempted TLC on Silica Gel with EtOAc caused hydrolysis of the salt to 2-furylboronic acid. Its IR (KBr) had bands at  $\nu_{\text{max}}$  1575 (C=C), 1005 (B-F), 970 (B-F)  $\text{cm}^{-1}$ , the  $^1\text{H}$  NMR ( $\text{CD}_3\text{CN}$ , 200MHz) had  $\delta$  at 7.44-7.33 (m, 1H), 6.25-

6.17 (m, 1H), 6.17-6.10 (m, 1H); and the  $^{11}\text{B}$  NMR ( $\text{CD}_3\text{CN}$ , 160MHz) has  $\delta$  (from  $\text{BF}_3\cdot\text{OEt}_2$ , 0 ppm) 1.8 (q,  $J$  49Hz) and it had the correct elemental analysis for C and H. It is air stable and has been used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Vedejs et al. *J Org Chem* **60** 3020 1995].

**Potassium isoamyl xanthate (potassium *O*-isoamyl dithiocarbonate)** [61792-26-5] **M 202.4,  $\text{pK}^{25}$  1.82 ( $\text{pK}^0$  2.8 free acid)**. Crystallise it twice from acetone/diethyl ether. Dry it in a desiccator for two days and store it under refrigeration. Its solubility in  $\text{Me}_2\text{CO}$  is 2.3% at 56°. [See potassium ethyl xanthate above for storage, Warren & Matthews *Anal Chem* **26** 1985 1954, *Beilstein* **3** III 340, **3** IV 404.]

**Potassium 4-methoxyphenyltrifluoroborate** [192863-36-8] **M 214.0,  $m >300^\circ$** . The salt is prepared by mixing 4-methoxyphenylboronic acid (0.65g, 4.3mmol) and  $\text{KHF}_2$  (0.76g, 9.7mmol) in  $\text{H}_2\text{O}$  (1ml) and MeOH (1.8ml) at  $\sim 25^\circ$  for 2 hours and the resulting yellow slurry is taken up in  $\text{Me}_2\text{CO}$  (10ml) and evaporated under a vacuum. The residue is dissolved consecutively in hot  $\text{Me}_2\text{CO}$  and THF, filtered and  $\text{Et}_2\text{O}$  is added to give a yellow solid which is filtered off, washed with  $\text{Et}_2\text{O}$  until free from the yellow colour, and the crystalline material is dried in a Schlenk vacuum line to give *p*- $\text{MeOC}_6\text{H}_4\text{-BF}_3^- \text{K}^+$  (0.83g, 92%). Its  $^1\text{H}$  NMR (acetone- $d_6$ , 500MHz) has  $\delta$  at 7.38 (d,  $J = 8.2\text{Hz}$ , 2H), 7.69 (d,  $J = 8.2\text{Hz}$ , 2H), 3.70 (s, 3H); the  $^{13}\text{C}$  NMR (acetone- $d_6$ , 125MHz) has  $\delta$  at 158.9, 133.5, 112.8, 55.1; the  $^{19}\text{F}$  NMR (acetone- $d_6$ , 470MHz) has  $\delta$  (from  $\text{CFCl}_3$ , 0ppm) at -142.28 (br d,  $J = 66\text{Hz}$ ); and the  $^{11}\text{B}$  NMR (acetone- $d_6$ , 64MHz) has  $\delta$  (from  $\text{BF}_3\cdot\text{OEt}_2$ , 0ppm) at 4.8 (br s,  $J = 17\text{Hz}$ ), and it has the correct elemental analysis for C and H. It is air stable and has been used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *Org Lett* **4** 1867 2002].

**Potassium 3-methoxyphenyltrifluoroborate** [438553-44-7] **M 214.0,  $m >300^\circ$**  was prepared in the same manner from 3-methoxyphenylboronic acid in 89% yield and had IR (KBr) with bands at  $\nu_{\text{max}}$  1241 (? C=C), 987 (B-F)  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (acetone- $d_6$ , N, 500MHz) had  $\delta$ : 7.03 (m, 3H), 6.61 (m, 1H), 3.67 (s, 3H); the  $^{13}\text{C}$  NMR (acetone- $d_6$ , 125MHz) had  $\delta$  at 159.6, 128.2, 124.9, 117.6, 112.0 55.0; the  $^{19}\text{F}$  NMR (acetone- $d_6$ , 470MHz) had  $\delta$  (from  $\text{CFCl}_3$ , 0ppm) at -142.8 (br d,  $J = 66\text{Hz}$ ); and the  $^{11}\text{B}$  NMR (acetone- $d_6$ , 64MHz) had  $\delta$  (from  $\text{BF}_3\cdot\text{OEt}_2$ , 0ppm) at 4.2 (br s,  $J$  43Hz), and it had the correct elemental analysis for C and H [Molander & Biolatto *Org Lett* **4** 1867 2002, cf Vedejs et al. *J Org Chem* **60** 3020 1995].

**Potassium 2-methoxyphenyltrifluoroborate** [236388-46-8] **M 214.0,  $m >300^\circ$**  can be prepared in the same manner from 2-methoxyphenylboronic acid and used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling reactions.

**Potassium methyltrifluoroborate** [13862-28-7] **M 121.94,  $m$  168-183°, 183°**. This salt, which was used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling is air stable, more robust and easier to handle than methylboronic acid. It is prepared by adding trimethyl boroxine (5.14g, 41mmol, 3 equivs, [823-96-1]) at room temperature to  $\text{KHF}_2$  (19.0g, 243mmol, 6 equivs) in MeCN (200ml), and the mixture is cooled to  $0^\circ$  and stirred for 30 minutes. Then  $\text{H}_2\text{O}$  (4.5ml) is added and after stirring for 3 hours the solvent is evaporated and the resulting solid is dried thoroughly *in vacuo*. The residue is triturated with  $\text{Me}_2\text{CO}/\text{MeOH}$  (1:1, 100ml), filtered off, washed with the same mixture of solvents (100ml) and once with  $\text{Me}_2\text{CO}/\text{MeOH}$  (1:2, 100ml), and the remaining insoluble white solid is dried at high vacuum in a Schlenk line to give  $\text{MeBF}_3^- \text{K}^+$  (12.0g, 80%) as a white powder  $m$  183°. Its  $^1\text{H}$  NMR ( $\text{D}_2\text{O}$ , 500MHz) has  $\delta$  at -0.15 (s), the  $^{13}\text{C}$  NMR ( $\text{D}_2\text{O}$ , 125MHz) has  $\delta$  at 1.40 to -1.14 (br s),  $^{19}\text{F}$  NMR ( $\text{D}_2\text{O}$ , 470MHz) has  $\delta$  at -132.3 (q,  $J = 64\text{Hz}$ ), and  $^{11}\text{B}$  NMR ( $\text{D}_2\text{O}$ , 64MHz) has  $\delta$  at 7.25 (q,  $J = 64\text{Hz}$ ). and has the correct elemental analysis for C and H. It has been used for inserting a methyl group in the position of a halogen or triflate group of various arenes. [Molander et al. *J Org Chem* **68** 5534 2003.]

**Potassium nonafluorobutane sulfonate** [29420-49-3] **M 338.2**. Wash it with  $\text{H}_2\text{O}$  and dry it *in vacuo*. When the K salt is distilled with 100%  $\text{H}_2\text{SO}_4$ , it gives the *free acid* which can be distilled (**b** 105°/22mm, 210-212°/760mm) and then converted to the pure K salt. [Gramstad & Haszeldine *J Chem Soc* 2640 1957, *Beilstein* **2** IV 818.]

**Potassium phenol-4-sulfonate (4-hydroxybenzene-1-sulfonic acid K salt)** [30145-40-5] **M 212.3**. Crystallise it several times from distilled water at  $90^\circ$ , after treatment with charcoal, and cooling to *ca*  $10^\circ$ . Dry it at  $90\text{-}100^\circ$ , *in vacuo*. [*Beilstein* **11** H 55, **11** I 242, **11** II 137, **11** III 498, **11** IV 582.]

**Potassium phenyltrifluoroborate** [329976-74-1] **M 212.1, m 290°, 296°(dec)**. The salt is obtained by adding excess of saturated aqueous  $\text{KHF}_2$  (125ml, ca 4.5M solution, 563mmol) dropwise to a solution of phenylboronic acid (20g, 169mmol, [98-80-6]) in MeOH (50ml) with vigorous stirring. The precipitate is collected after 15 minutes, washed with cold MeOH and recrystallised from the minimum volume of MeCN and dried *in vacuo* to give  $\text{Ph-BF}_3^- \text{K}^+$  (25.5g, 82%), m 296°(dec.). Its  $^1\text{H}$  NMR ( $\text{CD}_3\text{CN}$ , 200MHz) has  $\delta$  at 7.44-7.41 (m, 2H), 7.22-7.05 (m, 3H); the  $^{19}\text{F}$  NMR ( $\text{CD}_3\text{CN}$ , vs  $\text{CF}_3\text{C}_6\text{H}_5$ , 470MHz) has  $\delta$  at -79 (1:1:1:1 q,  $J = 57\text{Hz}$ ); and the  $^{11}\text{B}$  NMR ( $\text{CD}_3\text{CN}$ , 160MHz) has  $\delta$ : 4.1 (q,  $J = 57\text{Hz}$ ), and it has the correct elemental analysis for C and H. [Vedejs et al. *J Org Chem* **60** 3020 1995, Thierig & Umland *Naturwissenschaften* **54** 563 1967.] It is air stable and has been used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Vedejs et al. *J Org Chem* **60** 3020 1995, Molander & Biolatto *J Org Chem* **68** 4302 2003]. **Potassium 2-naphthalenetri-fluoroborate** [668984-08-5] **M 234.1, m >300°** can be prepared in the same manner from 2-naphthaleneboronic acid and used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling reactions.

**Potassium pyridine-3-trifluoroborate** [561328-69-6] **M 185.0, m 228-232°**. The 3-pyridine salt is prepared under  $\text{N}_2$  from 3-bromopyridine (3.46g, 21.9mmol) and triisopropyl borate (6.1ml, 26.4mmol) in a mixture of dry toluene (40ml) and dry THF (10ml) at  $-60^\circ$  to which is added *n*-BuLi (39ml, 1.6M in hexane, 62mmol) dropwise, the mixture is stirred for 2 hours and allowed to warm to room temperature overnight. This mixture containing the pyridine-3-boronic lithium salt is cooled to  $0^\circ$  and an aqueous solution of  $\text{KHF}_2$  (4.5M, 58mmol) is added dropwise and stirred for 4 hours then the mixture is evaporated to dryness. The residue is dissolved in MeOH, filtered, evaporated to dryness and the residual oil gave a solid (90%) under high vacuum. Its  $^1\text{H}$  NMR ( $\text{MeOD-}d_3$ , 500MHz) has  $\delta$  at 8.57 (br s, 1H), 8.26 (br d,  $J = 4.6\text{Hz}$ , 1H), 7.90 (d,  $J = 7.3$ , 1H), 7.23 (dd,  $J = 5.5$ , 6.6Hz, 1H); the  $^{13}\text{C}$  NMR ( $\text{DMSO-}d_6$ , 125MHz) has  $\delta$ : 153.0, 145.7, 138.8, 122.0; the  $^{19}\text{F}$  NMR ( $\text{DMSO-}d_6$ , 470MHz) has  $\delta$ : -134.9 (s); and the  $^{11}\text{B}$  NMR ( $\text{MeOD-}d_3$ , 64MHz) has  $\delta$ : 4.0 (br s). It has been used for inserting a 2,6-dimethylphenyl group into the position of a halogen or triflate group of various arenes and various heterocycles in Pd-catalysed Suzuki-Miyaura cross-coupling reactions [Molander & Biolatto *J Org Chem* **68** 4302 2003.]

**Potassium tetraphenylborate** [3244-41-5] **M 358.3**. Precipitate it from a solution of KCl acidified with dilute HCl, then crystallise it twice from acetone, wash it thoroughly with water and dry it at  $110^\circ$  [Findeis & de Vries *Anal Chem* **28** 1899 1956]. It has also been recrystallised from conductivity water. [Beilstein **16** IV 1625.]

**Potassium 3-thiophenyltrifluoroborate** [192863-37-9] **M 190.0, m >260°(dec)**. This salt, which was used successfully in Pd-catalysed Suzuki-Miyaura cross-coupling is air stable, more robust and easier to handle than 3-thienyl boronic acid. 3-Thiophenylboronic acid (4.97g, 26.2mmol) and  $\text{KHF}_2$  (5.14g, 65.8mmol) in a Nalgene (polyethylene) bottle (100ml) are stirred vigorously with MeOH (7.5ml) and  $\text{H}_2\text{O}$  (14ml) for 2 hours, and the amber solid formed is set aside at  $4^\circ$  for 2 hours. The solid was collected washed with the minimum of cold MeOH and dissolved in hot  $\text{Me}_2\text{CO}$  filtered, and the filtrate was cooled to  $25^\circ$  and  $\text{Et}_2\text{O}$  was added in portions with stirring, until the supernatant showed no cloudiness. The mixture was set aside at  $4^\circ$  for 1 hour until crystallisation was complete. The crystals were collected washed with a little cold  $\text{Et}_2\text{O}$  and dried *in vacuo* to give  $\text{C}_4\text{H}_3\text{S-BF}_3^- \text{K}^+$  (4.68g, 94%). Its  $^1\text{H}$  NMR [ $(\text{CD}_3)_3\text{CO}$ , 500MHz] has  $\delta$  at 7.20 (s, 1H), 7.14 (m, 2H); the  $^{13}\text{C}$  NMR [ $(\text{CD}_3)_3\text{CO}$ , 125MHz] has  $\delta$  at 131.8, 125.2, 122.3; the  $^{19}\text{F}$  NMR [ $(\text{CD}_3)_3\text{CO}$ , 470MHz] has  $\delta$  at -139.5 (d,  $J = 75\text{Hz}$ ), and it has the correct elemental analysis for C and H. It has been used for inserting a thiophenyl group into the position of a halogen or triflate group of various arenes and various heterocycles. [Molander & Biolatto *J Org Chem* **68** 4302 2003.] **Potassium 5-methyl-2-thiophenyltrifluoroborate** [871231-40-2] **M 204.1, m >220°, >250° (dec also reported)** can be similarly prepared from 5-methylthiophenyl-2-boronic acid and used in Pd catalysed Suzuki cross-coupling reactions. •

**Potassium trimethylsilanolate (trimethylsilanol K salt)** [10519-96-7] **M 128.3, m 131-135° (cubic form),  $d^{25}$  1.11, 125°dec (orthorhombic form)**. Recrystallise it from  $\text{H}_2\text{O}$  and dry it at  $100^\circ/1\text{-}2\text{mm}$ . [Hyde et al. *J Am Chem Soc* **75** 5615 1953, IR: Tatlock & Rochow *J Org Chem* **17** 1555 1952, Beilstein **4** IV 3992.]

**Propargyl triphenyl phosphonium bromide** [2091-46-5] **M 381.4, m 179°**. It recrystallises from 2-propanol as white plates. It also crystallises from EtOH with m 156-158°. Its IR has  $\nu_{\text{max}}$  at 1440,  $1110\text{cm}^{-1}$  (P-C str). [Elter & Dediger *Justus Liebigs Ann Chem* **682** 62 1965, Schweizer et al. *J Org Chem* **42** 200 1977].

**Propenyloxy trimethylsilane** [1833-53-0] **M 130.3, b 93-95°/atm, d<sub>4</sub><sup>20</sup> 0.786.** Purify it by fractional distillation using a very efficient column at atmospheric pressure. It usually contains 5% of hexamethyl-disiloxane that boils at 99-101°, but is generally non-reactive and need not be removed. [Hauser & Hance *J Am Chem Soc* **71** 5091 1952.] It has been distilled under N<sub>2</sub> through a 15cm column packed with glass helices. Fraction **b** 99-104° is further purified by gas chromatography through a Carbowax column (Autoprep A 700) at a column temperature of 87°, and has a retention time of 9.5 minutes. [Krüger & Rochow *J Organomet Chem* **1** 476 1963-4.]

**1-Propenyltrimethylsilane (cis and trans mixture)** [17680-01-2] **M 114.3, b 85-88°, n<sub>D</sub><sup>20</sup> 1.4121.** Dissolve ~20g in THF (200ml), shake it with H<sub>2</sub>O (2x 300 ml), dry (Na<sub>2</sub>SO<sub>4</sub>) and fractionate. This is a mixture of *cis* and *trans* isomers which can be separated by gas chromatography on an AgNO<sub>3</sub> column (for preparation: see Seyferth & Vaughan *J Organomet Chem* **1** 138 1963) at 25° with He as carrier gas at 9 psi. The *cis*-isomer has **n<sub>D</sub><sup>25</sup> 1.4105**, and the *trans*-isomer has **n<sub>D</sub><sup>25</sup> 1.4062**. [Seyferth et al. *Pure Appl Chem* **13** 159 1966.]

**Propylphosphonic acid (1-propanephosphonic acid)** [4672-38-2] **M 124.1, m 73°, pK<sub>1</sub><sup>25</sup> 2.49, pK<sub>2</sub><sup>25</sup> 8.18 (H<sub>2</sub>O).** The phosphonic acid is purified by recrystallisation from hexane, heptane or \*C<sub>6</sub>H<sub>6</sub> to give long colourless needles, and is dried *in vacuo* over KOH. It is best prepared from di-*n*-butylphosphite (50mmol), which is converted into its sodium salt in dry hexane (150ml, by stirring under reflux until Na has dissolved, ~3-5 hours), treated with an equivalent of *n*-propylbromide and refluxed gently for 5-6 hours. After cooling, the mixture is washed thoroughly with H<sub>2</sub>O, the organic layer is dried by distillation under a vacuum, and the residual dibutyl *n*-propylphosphonate is refluxed with 50-70ml of concentrated HCl overnight, and distilled from an oil bath to remove, BuCl and BuOH until *ca* 30ml is left, then carefully evaporated *in vacuo* and the residual solid is recrystallised from \*C<sub>6</sub>H<sub>6</sub> and/or hexane to give *n*-propylphosphonic acid (~80% yield) **m 72.5-74.5°**. **Diethyl propanephosphonate** [18812-51-6] **M 180.2** has **b 88-89°/9mm, d<sub>4</sub><sup>25</sup> 1.010, n<sub>D</sub><sup>20</sup> 1.4172**, and its <sup>13</sup>C NMR (25MHz, MeCO *d*<sub>6</sub>, TMS) has δ<sub>C</sub> at 61.8 (ester CH<sub>3</sub>, <sup>3</sup>J<sub>P,C-3</sub> = 5.9 Hz), 61.3 (ester CH<sub>2</sub>, <sup>2</sup>J<sub>P,C-2</sub> = 6.2 Hz), 28.2 (P-CH<sub>2</sub>, <sup>1</sup>J<sub>P,C-1</sub> = 140.4 Hz), 16.7 (-CH<sub>2</sub>-CH<sub>2</sub>-CH<sub>3</sub>, <sup>2</sup>J<sub>P,C-2</sub> = 5.2 Hz) and 15.4 (propane CH<sub>3</sub>, <sup>3</sup>J<sub>P,C-3</sub> = 16.2 Hz) ppm [Ernst *Org Mag Res* **9** 35 1977]. [Kosolapoff *J Am Chem Soc* **67** 1180 1945, *Beilstein* **4** H 596.]

**Propylphosphonic anhydride (2,4,6-tripropyl-1,3,5,2,4,6-trioxatriphosphorinane-2,4,6-trioxide, T3P)** [68957-94-8] **M 318.2 (cyclic trimer), b 200°/0.3mm 200-350°/0.01-50mm.** This reagent is prepared by heating *n*-propylphosphonic acid with acetic anhydride at 70-100°, and the polymeric phosphonic acid anhydride intermediate gives the trimeric cyclic anhydride on distilling at 200-350°/0.01-50mm. The anhydride is immediately made into 50% w/w solutions in DMF, CH<sub>2</sub>Cl<sub>2</sub>, EtOAc or BuOAc, which are also commercially available. It was originally used as a peptide coupling reagent, but has found many applications which require removal of the elements of water from organic molecules. It allows the synthesis of a variety of heterocyclic compounds, acylation reactions which involves C-C coupling, in the presence of DMSO alcohols can be oxidised to ketones; and with NH<sub>2</sub>OH.HCl and Et<sub>3</sub>N in MeCN, TBDMSCl is a selective reagent for hydroxyamidation of carboxylic acids. [Llanes Garcia *Synlett* 1328 2007, Wehner et al. PCT Int. Appl. WO 2005014604 2005, *Chem Abstr* **142** 198208 2005, Wissmann & Kleiner *Angew Chem Int Ed Engl* **19** 133 1980, Escher & Bünning *Angew Chem Int Ed Engl* **25** 277 1986.]

**3-(2-Pyridyl)-5,6-diphenyl-1,2,4-triazine-*p,p'*-disulfonic acid, monosodium salt (H<sub>2</sub>O) [Ferrozine]** [63451-29-6] **M 510.5, m >350°(dec).** Purify it by recrystallisation from water or by dissolving it in the minimum volume of water, followed by addition of EtOH to precipitate the pure salt. It is light sensitive and complexes with Fe. [Stookey *Anal Chem* **42** 779 1970.]

**Pyrocatechol Violet (tetraphenolictriphenylmethanesulfonic acid Na salt)** [115-41-3] **M 386.4, ε 1.4 x 10<sup>4</sup> at 445nm in acetate buffer pH 5.2-5.4, pK<sub>Est(1)</sub>>0 (SO<sub>3</sub>H), pK<sub>Est(2)</sub>~9.4, pK<sub>Est(3)</sub>~13.** It is recrystallised from glacial acetic acid. It is very *hygroscopic* and is an indicator standard for metal complex titrations. [Mustafin et al. *Zh Anal Khim* **22** 1808 1967, *Beilstein* **19/3** V 703.]

**Pyrogallol Red (tetraphenolic xanthyliumphenylsulfonate)** [32638-88-3] **M 418.4, m >300°(dec), ε 4.3 x 10<sup>4</sup> at 542nm, pH 7.9-8.6, pK<sub>1</sub> 2.71, pK<sub>2</sub> 6.60, pK<sub>3</sub> 10.41, pK<sub>4</sub> 12.16 (5% aqueous EtOH).** It is recrystallised from aqueous alkaline solution (Na<sub>2</sub>CO<sub>3</sub> or NaOH) by precipitation on acidification. Filter the dye off and dry

it in a vacuum. [Suk *Col Czech Chem Commun* **31** 3127 1966, *Beilstein* **19** H 407, **19** II 417, **19** III/IV 599, **19/10** V 226.]

**Rose Bengal** [Acid Red 94, 4,5,6,7-tetrachloro-2'.4',5',7'-tetraiodofluorescein di-Na or di-K salt] [*di-Na salt* 632-69-9] **M 1017.7 (di-Na salt)** [*di-K salt* 11121-48-5] **M 1049.8 (di-K salt)**. This biological stain can be purified by chromatography on silica TLC using a 35:65 mix of EtOH/acetone as eluent. [*Beilstein* **19** II 261, **19** III/IV 2926.]

**Selenopyronine** [85051-91-8] **M 365.8**,  $\lambda_{\max}$  571nm ( $\epsilon$  81,000). Purify it as the hydrochloride from hydrochloric acid [Fanghanel et al. *J Phys Chem* **91** 3700 1987]. [*Beilstein* **18** II 434.]

**Selenourea** [630-10-4] **M 123.0**, **m 200°(slow heating)**, **202-205°**, **205-207°(dec)**, **214-215°(dec)**, **235°(dec)**. Recrystallise it from the least volume of H<sub>2</sub>O using Norite (preferably under N<sub>2</sub>) to form colourless needles which are dried over P<sub>2</sub>O<sub>5</sub>. It is air and light sensitive. It slowly turns moderately dark on storage even below 0°. [King & Hlavacek *J Am Chem Soc* **73** 1864 1951, Dunbar & Painter *J Am Chem Soc* **69** 1833 1947, Bacher & Bos *Rec Trav Chim Pays Bas* **62** 580 1943, Hope *Acta Chem Scand* **18** 1800 1964.] The *Se-methyl iodide* provides yellow crystals from EtOH/Et<sub>2</sub>O with **m 187-188°(dec)**. The *N,N-dimethyl* derivative crystallises from H<sub>2</sub>O or EtOH as colourless needles which slowly turn pink, then gray on standing, and although slightly soluble in \*benzene it can be recrystallised from it and has **m 167-170°(dec)** [Zingaro et al. *J Org Chem* **18** 292 1953, IR: Jensen & Nielsen *Acta Chem Scand* **20** 597 1966, *Beilstein* **3** IV 435.]

**Silicon tetraacetate** [562-90-3] **M 264.3**, **m 110-111°**, **b 148°/5-6mm**, **pK<sub>1</sub><sup>25</sup> 9.7**, **pK<sub>2</sub><sup>25</sup> 11.9 (for H<sub>4</sub>SiO<sub>4</sub> free acid)**. It can be crystallised from mixtures of CCl<sub>4</sub> and petroleum ether or Et<sub>2</sub>O, or from acetic anhydride and then dried in a vacuum desiccator over KOH. Ac<sub>2</sub>O adheres to the crystals and is removed first by drying at room temperature, then at 100° for several hours. It is soluble in Me<sub>2</sub>CO, is very *hygroscopic* and effervesces with H<sub>2</sub>O. It decomposes at 160-170°. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 701 1963, *Beilstein* **2** H 171.]

**Silver diethyldithiocarbamate** [1470-61-7] **M 512.3**, **m 172°**, **174°**, **176-78°**, **pK<sub>1</sub><sup>25</sup> 3.36 (for *N,N*-diethyldithio-carbamate)**. Purify it by recrystallisation from pyridine or CO<sub>2</sub>. Store it in a desiccator in a cool and dark place. [*Beilstein* **4** III 224, **4** IV 391.]

**Silver tosylate** [16836-95-6] **M 279.1**. The anhydrous salt is obtained by recrystallisation from H<sub>2</sub>O. Store it in the dark. [Claesson & Wallin *Chem Ber* **12** 1851 1879, *Beilstein* **11** H 97, 99.]

**Silver trifluoromethanesulfonate** [2923-28-6] **M 256.9**. Recrystallise it twice from hot CCl<sub>4</sub> [Alo et al. *J Chem Soc, Perkin Trans 1* 805 1986]. Store it in the dark. [*Beilstein* **3** IV 34.]

**Sodium *n*-alkylsulfates**. Recrystallise these salts from EtOH/Me<sub>2</sub>CO [Hashimoto & Thomas *J Am Chem Soc* **107** 4655 1985].

**Sodium amylpenicillin** [575-47-3] **M 350.4**, **m 188°(dec, anhydrous)** [ $\alpha$ ]<sub>D</sub><sup>23</sup> +319° (c 1, H<sub>2</sub>O). The *monohydrate* crystallises from moist acetone or moist ethyl acetate. Dry it in a vacuum. [Wintersteiner in "The Chemistry of Penicillin", Clarke, Johnson and Robinson eds, Princeton University Press, p 470 1949.]

**Sodium 9,10-anthraquinone-1,5-disulfonate (H<sub>2</sub>O)** [853-35-0] **M 412.3**. Separate it from insoluble impurities by continuous extraction with water. Recrystallise it twice from hot water and dry it under vacuum. [*Beilstein* **11** II 195, **11** III 634.]

**Sodium 9,10-anthraquinone-1-sulfonate (H<sub>2</sub>O)** [107439-61-2] **M 328.3**. Recrystallise it from hot water (4ml/g) after treatment with active charcoal, or from water by addition of EtOH. Dry it under vacuum over CaCl<sub>2</sub>, or in an oven at 70°. Store it in the dark. [*Beilstein* **11** II 192, **11** III 626, **11** IV 670.]

**Sodium 9,10-anthraquinone-2-sulfonate (H<sub>2</sub>O) (9,10-anthraquinone-2-sulfonic acid [Na salt, H<sub>2</sub>O])** [131-08-8] **M 328.3, pK<sub>Est</sub> <0 (SO<sub>3</sub>H)**. It crystallises from H<sub>2</sub>O or MeOH (charcoal). [Costa and Bookfield *J Chem Soc, Faraday Trans 1* **82** 991 1986, *Beilstein* **11** IV 671.]

**Sodium benzenesulfinate (benzenesulfonic acid Na salt)** [873-55-2] **M 164.2, m >300°, pK<sub>25</sub> 2.16 (2.74, for PhSO<sub>2</sub>H)**. Dissolve it in the minimum volume of O<sub>2</sub> free H<sub>2</sub>O (prepared by bubbling N<sub>2</sub> through for 2 hours) and adding O<sub>2</sub> free EtOH (prepared as for H<sub>2</sub>O), set aside at 4° overnight under N<sub>2</sub>, filter, wash with EtOH, then Et<sub>2</sub>O and dry *in vacuo*. The Na salt is relatively stable to air oxidation, but is best kept under N<sub>2</sub> in the dark. Also recrystallise it from EtOH and dry it at 120° for 4 hours in a vacuum. [Kornblum & Wade *J Org Chem* **52** 5301 1987, *Beilstein* **11** II 2, **11** IV 3.]

**Sodium benzenesulfonate** [515-42-4] **M 150.1, pK<sub>1</sub><sup>25</sup> 0.70 (2.55) (for PhSO<sub>3</sub>H<sub>2</sub>)**. Crystallise it from EtOH or aqueous 70-100% MeOH, and dry it under a vacuum at 80-100°. [*Beilstein* **11** H 28, **11** I 10, **11** II 18, **11** III 33, **11** IV 27.]

**Sodium bis(trimethylsilyl)amide (hexamethyl disilazane sodium salt)** [1070-89-9] **M 183.4, m 165-167°(sintering at 140°)**. It can be sublimed at 170°/2mm (bath temperature 220-250°) onto a cold finger, and can be recrystallised from \*C<sub>6</sub>H<sub>6</sub> (its solubility is: 10g in 100ml at 60°). It is slightly soluble in Et<sub>2</sub>O and is decomposed by H<sub>2</sub>O. [Wannagat & Niederprüm *Chem Ber* **94** 1540 1961.] It is available commercially under N<sub>2</sub> in Sure/Seal bottles in tetrahydrofuran (various concentrations) and at ~0.6M in toluene. [*Beilstein* **4** IV 4014.]

**Sodium 4-bromobenzenesulfonate** [5015-75-8] **M 258.7**. Crystallise it from MeOH, EtOH or distilled water. [*Beilstein* **11** H 570, **11** I 14, **11** II 30, **11** III 97.]

**Sodium cacodylate (3H<sub>2</sub>O)** [124-65-2] **M 214.0, m 60°**. Recrystallise it from aqueous EtOH. [*Beilstein* **4** H 612, **4** I 576, **4** IV 1818, see cacodylic acid above.]

**Sodium 4-chlorobenzenesulfonate** [5138-90-9] **M 214.6, pK<sub>Est</sub> <0 (for SO<sub>3</sub>H)**. Crystallise it twice from MeOH and dry it under vacuum. [*Beilstein* **11** IV 107.]

**Sodium 3-chloro-5-methylbenzenesulfonate** [5138-92-1] **M 228.7, pK<sub>Est</sub> <0 (for SO<sub>3</sub>H)**. Crystallise it twice from MeOH and dry it under vacuum. [*Beilstein* **2** I 22.]

**Sodium *p*-cymenesulfonate** [77060-21-0] **M 236.3**. Dissolve the salt in water, filter and evaporate to dryness. Recrystallise it twice from absolute EtOH and dry it at 110°.

**Sodium 1-decanesulfonate** [13419-61-9] **M 244.33**. Recrystallise it from absolute EtOH and dry it over silica gel. [*Beilstein* **4** IV 62.]

**Sodium *n*-decylsulfate** [142-87-0] **M 239.3**. Rigorously purify it by continuous Et<sub>2</sub>O extraction of a 1% aqueous solution for two weeks. [*Beilstein* **1** IV 1818.]

**Sodium dibenzylthiocarbamate** [55310-46-8] **M 295.4, m 230°(dec), pK<sub>20</sub> 3.13 (for monobenzylthiocarbamic acid)**. The free acid, when recrystallised twice from dry Et<sub>2</sub>O, has **m** 80-82°. The Na salt is reprecipitated from aqueous EtOH or EtOH by addition of Et<sub>2</sub>O or Me<sub>2</sub>CO [Lindler *Anal Chem* **50** 896 1978]. The *NH<sub>4</sub>* salt has **m** 130-133°, the *Cu* salt (yellow crystals) has **m** 284-286°, and the *Ti* salt has **m** 64-70°. [*Beilstein* **12** IV 2275.]

**Sodium 2,5-dichlorobenzenesulfonate** [5138-93-2] **M 249.0, pK<sub>Est</sub> <0 (for SO<sub>3</sub>H)**. Crystallise it from MeOH, and dry it under vacuum. [*Beilstein* **11** III 93.]

**Sodium diethylthiocarbamate (3H<sub>2</sub>O)** [20624-25-3] **M 225.3, m 94-96°(anhydrous), 98°, pK<sub>20</sub> 3.65**

(diethyldithiocarbamic acid). Recrystallise it from water. [*Beilstein* 4 III 224, 4 IV 390.]

**Sodium di(ethylhexyl)sulfosuccinate (Aerosol-OT, sodium docosate)** [577-11-7] **M 444.6**. Dissolve it in MeOH and the inorganic salts which precipitate are filtered off. Water is added and the solution is extracted several times with hexane. The residue is evaporated to one-fifth its original volume, \*benzene is added and azeotropic distillation is continued until no water remains. The solvent is evaporated. The white residual solid is crushed and dried *in vacuo* over P<sub>2</sub>O<sub>5</sub> for 48 hours [El Seoud & Fendler *J Chem Soc, Faraday Trans 1* **71** 452 1975]. [*Beilstein* 4 IV 114.] It solubilises major myelin trans membrane proteolipids, and forms reverse micelles in hydrocarbon solvents.

**Sodium 2,2'-dihydroxy-1-naphthaleneazobenzene-5'-sulfonate**. See Solochrome Violet R [2092-55-9] in "Aromatic Compounds" in this Chapter.

**Sodium 2,4-dihydroxyphenylazobenzene-4'-sulfonate (Tropaeolin O, Acid Orange 6)** [547-57-9] **M 316.3**. Recrystallise it from absolute EtOH. It has  $\lambda_{\max}$  at 490nm and CI (colour index) 14270. [*Beilstein* 16 H 275.]

**Sodium *p*-(*p*-dimethylaminobenzeneazo)-benzenesulfonate** [23398-40-5] **M 327.3**. Recrystallise it from water and dry it *in vacuo*. [*Beilstein* 16 H 331, 16 I 317, 16 II 169, 16 III 370, 16 IV 510.]

**Sodium 2,4-dimethylbenzenesulfonate** [827-21-4] **M 208.2**. Crystallise it from MeOH and dry it under vacuum. [*Beilstein* 11 III 339.]

**Sodium 2,5-dimethylbenzenesulfonate (sodium *p*-xylenesulfonate)** [827-19-0] **M 208.2**. Dissolve it in distilled water, filter it, then evaporate it to dryness. Recrystallise it twice from absolute EtOH or MeOH and dry it at 110° under vacuum. [*Beilstein* 11 III 93, 11 IV 502.]

**Sodium dimethyldithiocarbamate hydrate** [128-04-1 (*hydrate*), 207233-95-2 (*anhydrous*)] **M 143.2, m 106-108°, 120-122°, pK<sup>25</sup> 3.36 (diethyldithiocarbamic acid)**. Crystallise it from a small volume of H<sub>2</sub>O, or dissolve it in the minimum volume of H<sub>2</sub>O and add cold Me<sub>2</sub>CO, collect it and dry it in air. The solubility in Me<sub>2</sub>CO is ~12.5g/100ml. The *dihydrate* loses H<sub>2</sub>O on heating at 115° to give the *hemi-hydrate* that decomposes on further heating [Kulka *Can J Chem* **34** 1096 1956]. [*Beilstein* 4 IV 233.]

**Sodium *N,N*-dimethylsulfanilate** [2244-40-8] **M 223.2, m >300°**. It crystallises from water. [See *Beilstein* 4 IV 270.]

**Sodium 1-dodecanesulfonate** [2386-53-0] **M 272.4, m >300°**. Recrystallise it twice from EtOH and dry it in an oven at 105° for 2 hours. It picks up moisture to form the 3.5H<sub>2</sub>O. Its *hydrate* crystallises in several phases. [Tartar & Wright *J Am Chem Soc* **61** 543 1939, Reed & Tartar *J Am Chem Soc I* **57** 571 1935, *Beilstein* 4 III 27, 4 IV 64.]

**Sodium 4-dodecylbenzenesulfonate** [25155-30-0] **M 348.5**. It crystallises from propan-2-ol or H<sub>2</sub>O. [Gray et al. *J Org Chem* **20** 515 1955, *Beilstein* 11 IV 514.]

**Sodium dodecylsulfate (SDS, sodium laurylsulfate)** [151-21-3] **M 288.4, m 204-207°**. Purify this detergent by Soxhlet extraction with petroleum ether for 24 hours, followed by dissolution in acetone/MeOH/H<sub>2</sub>O 90:5:5(v/v) and recrystallisation [Politi et al. *J Phys Chem* **89** 2345 1985]. It has been purified by two recrystallisations from absolute EtOH, aqueous 95% EtOH, MeOH, isopropanol or a 1:1 mixture of EtOH/isopropanol to remove dodecanol, and dried under vacuum [Ramesh & Labes *J Am Chem Soc* **109** 3228 1987]. SDS has also been purified by repeatedly foaming whereby a 0.15% aqueous solution is made to foam and the foam is discarded, then the H<sub>2</sub>O is removed *in vacuo* and the residue is diluted to the required concentrations [see Cockbain & McMullen *Trans Faraday Soc* **47** 322 1951], or by liquid-liquid extraction [see Harrold *J Colloid Sci* **15** 280 1960]. Dry it over silica gel. For DNA work it should be dissolved in excess MeOH passed through an activated charcoal column and evaporated until it crystallises out. It has also been purified by dissolving in hot 95% EtOH (14ml/g), filtering and cooling, then drying in a vacuum

desiccator. *Alternatively*, it is crystallised from H<sub>2</sub>O, vacuum dried, washed with anhydrous Et<sub>2</sub>O and dried in vacuum again. These operations are repeated five times [Maritato *J Phys Chem* **89** 1341 1985, Lennox and McClelland *J Am Chem Soc* **108** 3771 1986, Dressik *J Am Chem Soc* **108** 7567 1986]. [*Beilstein* **1** IV 1847.]

**Sodium ethylmercurithiosalicylate (Thimerosal)** [54-64-8] **M 404.8, m ~230°**. Recrystallise this antibacterial from EtOH/Et<sub>2</sub>O. **HIGHLY TOXIC**. [Trikojus *Nature* **158** 472 1940, *Beilstein* **10** III 213.]

**Sodium ethylsulfate** [546-74-7] **M 166.1**. Recrystallise it three times from MeOH/Et<sub>2</sub>O and dry it in a vacuum. [*Beilstein* **1** H 326, **1** I 164, **1** III 1317, **1** IV 1325.]

**Sodium formaldehyde sulfoxylate dihydrate (sodium hydroxymethylsulfinate, Rongalite)** [149-44-0] **M 134.1, m 63-64° (dihydrate)**. It crystallises from H<sub>2</sub>O as the *dihydrate* and decomposes at higher temperatures. Store it in a closed container in a cool place. It is insoluble in EtOH and Et<sub>2</sub>O and is a good reducing agent. [X-ray structure: Tuter *J Chem Soc* 3064 1955.] Note that this compound {HOCH<sub>2</sub>SO<sub>2</sub>Na} should not be confused with formaldehyde sodium bisulfite adduct {HOCH<sub>2</sub>SO<sub>3</sub>Na} from which it is prepared by reduction with Zn. [*Beilstein* **1** IV 3052.]

**Sodium hexadecylsulfate** [1120-01-0] **M 344.5**. Recrystallise it from absolute EtOH or MeOH and dry it in vacuum [Abu Hamdiyyah & Rahman *J Phys Chem* **91** 1531 1987].

**Sodium 2-hydroxy-4-methoxybenzophenone-5-sulfonate** [6628-37-1] **M 330.3**. Crystallise it from MeOH and dry it under vacuum.

**Sodium *p*-hydroxyphenylazobenzene-*p'*-sulfonate** [2623-36-1] **M 288.2**. Recrystallise it from 95% EtOH.

**Sodium 2-mercaptoethanesulfonate (MESNA)** [19767-45-4] **M 164.2, pK<sub>1</sub><sup>20</sup> <0 (SO<sub>3</sub><sup>-</sup>), pK<sub>2</sub><sup>20</sup> 9.53 (SH)**. It can be recrystallised from H<sub>2</sub>O and does not melt below 250°. It is purified further by converting to the free acid by passing a 2M solution through an ion-exchange (Amberlite IR-120) column in the acid form, evaporating the eluate in a vacuum to give the acid as a viscous oil (readily decomposes) which can be checked by acid and SH titration. It is then dissolved in H<sub>2</sub>O, carefully neutralised with aqueous NaOH, evaporated and the salt recrystallised from H<sub>2</sub>O [Schramm *J Am Chem Soc* **77** 6231 1955]. [*Beilstein* **4** IV 85.]

**Sodium metanilate** [1126-34-7] **M 195.2**. It crystallises from hot water. [*Beilstein* **14** H 688.]

**Sodium methanethiolate [sodium methylmercaptide]** [5188-07-8] **M 70.1, pK<sup>25</sup> 10.33 (MeS<sup>-</sup>)**. Dissolve the salt (10g) in EtOH (10ml) and add Et<sub>2</sub>O (100ml). Cool and collect the precipitate, wash it with Et<sub>2</sub>O and dry it in a vacuum. It is a white powder which is very soluble in EtOH and H<sub>2</sub>O. [Billmann & Jensen *Bull Soc Chim Fr* **3** 2318 1936, *Beilstein* **1** III 1212.]

**Sodium 3-methyl-1-butanesulfonate** [5343-41-9] **M 174.1**. It crystallises from 90% MeOH.

**Sodium 1-naphthalenesulfonate** [130-14-3] **M 230.2**. Recrystallise it from water or aqueous acetone [Okadata et al. *J Am Chem Soc* **108** 2863 1986]. [*Beilstein* **11** IV 521.]

**Sodium 2-naphthalenesulfonate** [532-02-5] **M 230.2**. It crystallises from hot 10% aqueous NaOH or water and is dried in a steam oven. [*Beilstein* **11** IV 521.]

**Sodium 2-naphthylamine-5,7-disulfonate (Amido-G-acid)** [79004-97-0] **M 235.4**. Crystallise it from water (charcoal) and dry it in a steam oven. [*Beilstein* **14** H 784, **14** II 473, **14** IV 2811.]

**Sodium 1-octanesulfonate H<sub>2</sub>O** [5324-84-5] **M 216.2**. Recrystallise it from absolute EtOH. [*Beilstein* **4** IV 58.]

**Sodium phenol-4-sulfonate (2H<sub>2</sub>O) (4-hydroxybenzenesulfonic acid Na salt)** [825-90-1] **M 232.2**. It crystallises from hot water (g/1m) by cooling to 0°, or from MeOH, and is dried in vacuum. [Beilstein 11 II 134.]

**Sodium piperazine-*N,N'*-bis(2-ethanesulfonate) H<sub>2</sub>O (PIPES-Na salt)** [76836-02-7] **M 364.3**. It crystallises from water and EtOH. [Beilstein 23/2 V 380.]

**Sodium isopropyl xanthate (sodium *O*-isopropylidithiocarbonate)** [140-93-2] **M 158.2, pK<sup>25</sup> 2.16 (for -S<sup>-</sup>)**. It crystallises from ligroin/ethanol.

**Sodium sulfanilate (sodium *p*-aminobenzenesulfonic acid)** [515-74-2] **M 195.2**. It crystallises from water. [Beilstein 14 IV 2655.]

**Sodium taurocholate [2-(3 $\alpha$ ,7 $\alpha$ ,12 $\alpha$ -trihydroxy-24-oxo-5- $\beta$ -cholan-24-ylamino)ethanesulfonic acid sodium salt monohydrate]** [145-42-6; 312693-83-7; 345909-26-4 (*x* H<sub>2</sub>O)] **M 555.7 (monohydrate), m 168°dec (hydrate), [ $\alpha$ ]<sub>D</sub><sup>20</sup> +23.9° (c 2.5, H<sub>2</sub>O), pK of acid is 1.4**. The non-sulfated bile salt has been synthesised from *ethyl cholate* (**m 162-163°**, crystallised from EtOAc/petroleum ether b 30-60° 2:8) *via* the *hydrazide* (**m 210°**, sintering at 202°), which was diazotised to the *azide* (NaNO<sub>2</sub>/HCl at 0-2°) and condensed with taurine in aqueous N NaOH at 8-14°/45 minutes. The acidified product was converted to *Na taurocholate* which was prepared and purified by precipitation with saturated aqueous NaCl and Et<sub>2</sub>O (84% recovery; note that crystallisation does not occur unless enough H<sub>2</sub>O is present) [Cortese *J Am Chem Soc* 59 2532 1937]. It was also purified by recrystallisation from aqueous EtOH/Et<sub>2</sub>O, or by gel chromatography using Sephadex LH-20. It is a useful anionic detergent for solubilising proteins and bilirubin [Wosiewicz & Schroebler *Experientia* 35 717 1979]. It has a CMC of 3-11 mM at 20-25°, with an average micellar weight of 2100. It is hydrolysed by mineral acids to cholic acid and taurine. [Tanaka *Z physiol Chem* 220 39 1933, Beilstein 10 III 1655, 10 IV 2078.]

**Sodium tetradecylsulfate (sodium myristyl sulfate)** [1191-50-0] **M 316.4**. It recrystallises from absolute EtOH [Abu Hamdiyyah & Rahman *J Phys Chem* 91 1531 1987]. It is *hygroscopic*. [Beilstein 1 H 716, 1 IV 1866.]

**Sodium tetrakis-(4-fluorophenyl)borate hydrate** [207683-22-5] **M 450.2**. This gravimetric reagent for Cs is purified by passing a solution (10g in 100ml H<sub>2</sub>O) through a column of Dowex 50Wx4 (Na form) and eluting with dilute NaCl. Extract the 250-275ml eluate with Et<sub>2</sub>O (3 x 50ml), add dry xylene (200ml), evaporate the Et<sub>2</sub>O off *in vacuo*, immerse the xylene in a bath at 50° and the salt crystallises out. It is *hygroscopic*. [Moore et al. *Anal Chim Acta* 35 1 1966, Tsubouci et al. *Anal Chem* 57 783 1985.]

**Sodium tetraphenylborate [tetraphenyl boron Na]** [143-66-8] **M 342.2**. Dissolve the borate in dry MeOH and add dry Et<sub>2</sub>O. Collect the solid and dry it in a vacuum at 80°/2mm for 4 hours. It can also be extracted (Soxhlet) using CHCl<sub>3</sub>, and it crystallises from CHCl<sub>3</sub> as snow-white needles. It is freely soluble in H<sub>2</sub>O, Me<sub>2</sub>CO but insoluble in petroleum ether and Et<sub>2</sub>O. An aqueous solution has pH ~5 and can be stored for days at 25° or lower, and for 5 days at 45° without deterioration. Its solubility in polar solvents increases with decrease in temperature [Wittig & Raff *Justus Liebigs Ann Chem* 573 204 1950]. The salt can also be recrystallised from acetone/hexane or CHCl<sub>3</sub>, or from Et<sub>2</sub>O/cyclohexane (3:2) by warming the solution to precipitate the compound. Dry it in a vacuum at 80°. It dissolves in Me<sub>2</sub>CO at 50-60° to give a clear solution. After standing at this temperature for 10 minutes the mixture is filtered rapidly through a pre-heated Büchner funnel, cooled and the crystals are collected and dried in a vacuum desiccator at room temperature for 3 days [Abraham et al. *J Chem Soc, Faraday Trans 1* 80 489 1984]. If the product gives a turbid aqueous solution, the turbidity can be removed by treating with freshly prepared alumina gel and filtering. [Beilstein 16 IV 1624.]

**Sodium thioglycolate** [367-51-1] **M 114.1**. It crystallises from 60% EtOH (charcoal). It is *hygroscopic*. [Beilstein 3 IV 600.]

**Sodium 4-toluenesulfinate** [824-79-3] **M 178.2, pK<sup>25</sup> 2.80 (1.99)(for -SO<sub>2</sub><sup>-</sup>)**. Recrystallise the salt from

water (to constant UV spectrum) and dry it under vacuum, or extract it with hot benzene, then dissolve it in EtOH/H<sub>2</sub>O and heat with decolorising charcoal. The solution is filtered and cooled to give crystals of the dihydrate. [Beilstein 11 I 718.]

**Sodium 4-toluenesulfonate** [657-84-1] **M 194.2, pK<sup>25</sup> -1.34 (for -SO<sub>3</sub><sup>-</sup>)**. Dissolve it in distilled water, filter it to remove insoluble impurities and evaporate it to dryness. Then recrystallise it from MeOH or EtOH, and dry it at 110°. Its solubility in EtOH is not high (maximum 2.5%), so that Soxhlet extraction with EtOH may be preferable. Sodium *p*-toluenesulfonate has also been crystallised from Et<sub>2</sub>O and dried under a vacuum at 50°. [Beilstein 11 I 4, 11 II 6, cf Gibson & Reid *J Chem Soc* 879 1923.]

**Sodium 2,2',4-trihydroxyazobenzene-5'-sulfonate** [3564-26-9] **M 295.3**. Purify the dye by precipitating the free acid from aqueous solution using concentrated HCl, then wash it and extract it with EtOH in a Soxhlet extractor. Evaporation of the EtOH leaves the purified acid which is converted to the sodium salt with an equivalent of NaOH or Na<sub>2</sub>CO<sub>3</sub>.

**Sodium 2,4,6-trimethylbenzenesulfonate** [6148-75-0] **M 222.1**. Crystallise it twice from MeOH and dry it under vacuum. [Beilstein 11 III 345.]

**Sodium trimethylsilanolate (sodium trimethylsilanol)** [18027-10-6] **M 112.2, m 230°(dec)**. It is very soluble in Et<sub>2</sub>O and C<sub>6</sub>H<sub>6</sub> but moderately soluble in petroleum ether. It is purified by sublimation at 130-150° in a high vacuum. [Hyde et al. *J Am Chem Soc* 75 5615 1953, Tatlock & Rochow *J Org Chem* 17 1555 1952, Beilstein 4 III 1856.]

**Sodium 3,5-xylenesulfonate** [30587-85-0] **M 208.2**. Dissolve it in distilled water, filter, then evaporate it to dryness and recrystallise it twice from absolute EtOH and then dry it at 110°. [Beilstein 11 H 126, 11 I 34.]

**Strontium thiosalicylate (5H<sub>2</sub>O)** [15123-90-7] **M 289.8**. It crystallises from hot water (0.5g/ml) by cooling to 0°. [Beilstein 10 IV 272.]

**Sulfur trioxide pyridine complex** [26412-87-3] **M 159.2, m 155-165°, 175°**. Wash the solid with a little CCl<sub>4</sub>, then H<sub>2</sub>O to remove traces of pyridine sulfate, and dry it over P<sub>2</sub>O<sub>5</sub>. [Baumgarten *Chem Ber* 59 1166 1926, Olah et al. *Synthesis* 59 1979, Beilstein 20/5 V 184.]

**Tetrabutylammonium borohydride** [33725-74-5] **M 257.3, m 128-129°**. Purify it by recrystallisation from EtOAc followed by careful drying under vacuum at 50-60°. Samples purified in this way showed no signs of loss of active H after storage at room temperature for more than 1 year. Nevertheless samples should be stored at ca 6° in tightly stoppered bottles if they are to be kept for long periods. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>. [Raber & Guida *J Org Chem* 41 690 1976, Brändström et al. *Tetrahedron Lett* 3173 1972.]

**Tetrabutylammonium tetrafluoroborate** [429-42-5] **M 329.3, m 161.8°, 161-163°, pK<sup>25</sup> -4.9 (for HBF<sub>4</sub>)**. Recrystallise it from H<sub>2</sub>O, aqueous EtOH or from EtOAc by cooling in Dry-ice. Also recrystallise it from ethyl acetate/pentane or dry acetonitrile. Dry it at 80° under vacuum. [Detty & Jones *J Am Chem Soc* 109 5666 1987, Hartley & Faulkner *J Am Chem Soc* 107 3436 1985.] The acetate has m 118±2° (from BuCl), the bromide has m 118° (from EtOAc) and the nitrate has m 120° (from C<sub>6</sub>H<sub>6</sub>). [Witschonka & Kraus *J Am Chem Soc* 69 2472 1947, Wheeler & Sandstedt *J Am Chem Soc* 77 2024 1955, Beilstein 4 IV 558.]

**Tetraethoxysilane (tetraethyl orthosilicate)** [78-10-4] **M 208.3, m -77°, b 165-166°/atm, d<sub>4</sub><sup>20</sup> 0.933, n<sub>D</sub><sup>25</sup> 1.382**. Fractionate it through an 80cm Podbielniak type column with a heated jacket and partial take-off head. It is soluble in EtOH, and is slowly decomposed by H<sub>2</sub>O and. It is **flammable**—it irritates the eyes and mucous membranes. [Sumrell & Ham *J Am Chem Soc* 78 5573 1956, Bradley et al. *J Chem Soc* 5020 1952, Beilstein 1 IV 1360.]

**Tetraethylammonium hexafluorophosphate** [429-07-2] **M 275.2, m >300°, 331°(dec), pK<sub>1</sub><sup>25</sup> ~ 0.5, pK<sub>2</sub><sup>25</sup>**

**5.12 (for fluorophosphoric acid  $\text{H}_2\text{PO}_3\text{F}$ ).** Dissolve the salt (0.8g) in hot  $\text{H}_2\text{O}$  (3.3ml) and cool to crystallise. Yield of prisms is 0.5g. Its solubility in  $\text{H}_2\text{O}$  is 8.1g/L at  $19^\circ$  [Lange & Müller *Chem Ber* **63** 1067 1930]. [Beilstein **4** III 199.]

**Tetraethylammonium tetrafluoroborate** [429-06-1] **M 217.1, m 235°, 356-367°, 275-277°, 289-291°.  $\text{pK}^{25}_{\text{a}} -4.9$  (for  $\text{HBF}_4$ ).** Dissolve the salt in hot MeOH, filter and add  $\text{Et}_2\text{O}$ . It is soluble in ethylene chloride [Thompson & Kraus *J Am Chem Soc* **69** 1016 1947, Wheeler & Sandstadt *J Am Chem Soc* **77** 2025 1955]. It has also been recrystallised three times from a mixture of ethyl acetate/hexane (5:1) or MeOH/petroleum ether, then stored at  $95^\circ$  for 48 hours under vacuum [Henry & Faulkner *J Am Chem Soc* **107** 3436 1985, Huang et al. *Anal Chem* **58** 2889 1986]. It is used as a supporting electrolyte. [Beilstein **4** IV 333.]

**Tetraethylsilane** [631-36-7] **M 144.3, b 153.8°/760mm,  $d_4^{30}$  0.77,  $n_D^{30}$  1.427.** Fractionate it through a 3ft vacuum-jacketed column packed with 1/4" stainless steel saddles. The material is finally percolated through a 2ft column packed with alumina and maintained in an inert atmosphere. [Staveley et al. *J Chem Soc* 1992 1954, Altshaller & Rosenblum *J Am Chem Soc* **77** 272 1955, Beilstein **4** H 625.]

**1.1.3.3-Tetraisopropyldisiloxane** [18043-71-5] **M 246.5, b 129-130°/6mm,  $d_4^{30}$  0.89,  $n_D^{30}$  1.47.** Fractionate it under reduced pressure in a  $\text{N}_2$  atmosphere. [Gilman & Clark *J Am Chem Soc* **69** 1500 1947.]

**Tetrakis(hydroxymethyl)phosphonium chloride (THPC)** [124-64-1] **M 190.6, m 151°.** THPC is prepared in an efficient fume cupboard by placing a mixture of 40% formaldehyde (90ml) and concentrated aqueous HCl (40ml, sp. gr. 1.2) into a 100ml 3-bulb Ladenburg flask (filling about half of the neck of the flask), the latter is provided with a stopper through which is passed a gas delivery tube that reaches the bottom. The flask is inclined to an almost horizontal position, when  $\text{PH}_3$  (from a cylinder) is passed to displace the air, and the solution is heated to  $\sim 80^\circ$  [note that at lower temperatures the reaction is slower, and at higher temperatures the vapour pressure of the solution decreases gas absorption]. The bulbs of the flask help to hold back the gas which results in a more complete reaction. The complete reaction may take up to many hours. Some  $\text{H}_2$  and  $\text{PH}_3$  which escape from the side neck of the flask should be vented through the flue of the fume cupboard. The mixture is then evaporated with stirring on a steam bath or *in vacuo* ( $\sim 50$  to  $60^\circ$ ) until a white granular solid is obtained (50g  $\sim$  88%). Its purity is good enough for most purposes. However, it can be purified to analytical purity by recrystallising 1g of THPC from 50ml of AcOH. The low melting flat needles obtained which contain AcOH of crystallisation are converted to the pure substance, **m 151°**, by blowing dry air at  $100^\circ$  over the crystals, or by drying them at  $100^\circ$  *in vacuo* to give a constant weight. It can also be recrystallised from absolute EtOH. [For a Kg scale laboratory preparation see Reeves et al. *J Am Chem Soc* **77** 3923 1955.] It is very deliquescent and should be kept in the presence of a desiccant. It is readily soluble in MeOH, slightly soluble in  $\text{CHCl}_3$ , but insoluble in  $\text{Et}_2\text{O}$ . It is quite stable, its aqueous solutions can be boiled without decomposition, and it is unaffected by dilute acids. The  $^1\text{H}$  NMR (60MHz,  $\text{D}_2\text{O}$ , DSS) has  $\delta$  for  $4\text{CH}_2$  at 4.77 (s,  $J_{\text{PCH}} = 1.7\text{Hz}$  and  $J_{\text{HCH}} \sim 153\text{Hz}$ ), and the  $^{31}\text{P}$  NMR (40MHz,  $\text{D}_2\text{O}$ , external 85%  $\text{H}_3\text{PO}_4$ ) has  $\delta$  at -25.8. With the stoichiometric amount of alkali, the free *tris(hydroxymethyl)phosphine* (see [2767-80-8] above) is obtained as an oil together with formaldehyde. However with excess of alkali on THPC,  $\text{H}_2$  evolution occurs with the liberation of HCHO and formation of *tris(hydroxymethyl)phosphine oxide* (see [1067-12-5] below) **m 54-55°**. The reaction with alkali becomes more complicated when the mixture is heated, and *bis(hydroxymethyl)phosphinic acid* is formed. An 80% w/v aqueous solution of THPC with  $d_4^{20}$  1.33 is available commercially. Also a 70-75% aqueous solution of *bis[tetrakis(hydroxymethyl)phosphonium] sulfate* [55566-30-8] **M 406.3** is commercially available. [Hoffman *J Am Chem Soc* **43** 1684 1921, Hoffman *J Am Chem Soc* **52** 2995 1930, Beilstein **1** IV 3062.]

**Tetramethoxysilane (tetramethyl orthosilicate)** [681-84-5] **M 152.2, m 4.5°, b 122°/760mm.** Purification is as for tetraethoxysilane. It has a vapour pressure of 2.5mm at  $0^\circ$ . [IR: Sternbach & MacDiarmid *J Am Chem Soc* **81** 5109 1959, Beilstein **1** IV 1266.]

**Tetramethylammonium borohydride** [16883-45-7] **M 89.0.** Recrystallisation of the borohydride from  $\text{H}_2\text{O}$  three times yields *ca* 94% pure compound. Dry in high vacuum at  $100^\circ$  for 3 hours. The solubility in  $\text{H}_2\text{O}$  is 48% ( $20^\circ$ ), 61% ( $40^\circ$ ), in EtOH it is 0.5% ( $25^\circ$ ), and in MeCN it is 0.4% ( $25^\circ$ ). It decomposes slowly in a

vacuum at 150°, but rapidly at 250°. The rate of hydrolysis of Me<sub>4</sub>N.BH<sub>4</sub> (5.8M) in H<sub>2</sub>O at 40° is constant over a period of 100 hours at 0.04% of original wt/hour. The rate decreases to 0.02%/hour in the presence of Me<sub>4</sub>NOH (5% of the wt of Me<sub>4</sub>N.BH<sub>4</sub>). [Banus et al. *J Am Chem Soc* **74** 2346 1952, *Beilstein* **4** IV 148.]

**Tetramethylammonium hexafluorophosphate** [558-32-7] **M 219.1, m >300°, d<sub>4</sub><sup>25</sup> 1.617, pK<sub>1</sub><sup>25</sup> ~ 0.5, pK<sub>2</sub><sup>25</sup> 5.12 (for fluorophosphoric acid H<sub>2</sub>PO<sub>3</sub>F).** The salt (0.63g) is recrystallised from boiling H<sub>2</sub>O (76ml), yielding pure (0.45) Me<sub>4</sub>N.PF<sub>6</sub> after drying at 100°. It is a good supporting electrolyte. [Lange & Müller *Chem Ber* **63** 1067 1930, *Beilstein* **4** III 110.]

**Tetramethylammonium triacetoxyborohydride** [109704-53-2] **M 263.1, m 93-98°, 96.5-98°.** If impure, wash it with freshly distilled Et<sub>2</sub>O and dry it overnight in a vacuum to give a free flowing powder. Check <sup>1</sup>H NMR, and if still suspect prepare it freshly from Me<sub>4</sub>NBH<sub>4</sub> and AcOH in \*C<sub>6</sub>H<sub>6</sub> and store it away from moisture [Banus et al. *J Am Chem Soc* **74** 2346 1952, Evans & Chipman *Tetrahedron Lett* **27** 5939 1986]. It is an **IRRITANT** and **MOISTURE SENSITIVE**.

**Tetramethylammonium triphenylborofluoride** [437-11-6] **M 392.2.** Crystallise it from acetone or acetone/ethanol. [*Beilstein* **4** IV 15.]

**2,4,6,8-Tetramethylcyclotetrasiloxane (TMCTS)** [2370-88-9] **M 240.5, m -69°, b 134°/750mm, 134.5-134.9°/755mm, d<sub>4</sub><sup>20</sup> 0.99, n<sub>D</sub><sup>20</sup> 1.3872.** It is purified by repeated redistillation, and fractions with the required <sup>1</sup>H NMR data are collected. [Sokolov *J Gen Chem USSR (Engl Transl)* **29** 262 1959, Sauer et al. *J Am Chem Soc* **68** 962 1946]. [*Beilstein* **4** IV 4099.]

**1,1,3,3-Tetramethyldisiloxane** [3277-26-7] **M 134.3, b 70.5-71°/731mm, 71-72°/atm, d<sub>4</sub><sup>30</sup> 0.75, n<sub>D</sub><sup>25</sup> 11.367.** Possible impurity is 1,1,5,5-tetramethyl-3-trimethylsiloxytrisiloxane **b 154-155°/733mm.** Fractionate it, collect fractions boiling below 80° and re-fractionate it. Its purity can be analysed by alkaline hydrolysis and measuring the volume of H<sub>2</sub> liberated followed by gravimetric estimation of silica in the hydrolysate. It is unchanged when stored in glass containers in the absence of moisture for 2-3 weeks. Small amounts of H<sub>2</sub> are liberated on long storage. *Care should be taken when opening a container due to developed pressure.* [Speier et al. *J Am Chem Soc* **79** 974 1958, Emeléus & Smythe *J Chem Soc* 609 1958, IR: Kriegsmann *Z Anorg Chem* **299** 78 1959, *Beilstein* **4** IV 3991.]

**N,N,N',N'-Tetramethylphosphonic diamide (methylphosphonic bis-dimethylamide)** [2511-17-3] **M 150.2, b 60.5°/0.6mm, 138°/32mm, 230-230°/atm, d<sub>4</sub><sup>30</sup> 1.0157, n<sub>D</sub><sup>30</sup> 1.4539.** Dissolve it in heptane or ethylbenzene, shake this with 30% aqueous NaOH, stir for 1 hour, separate the organic layer and fractionate. [Kosolapoff & Payne *J Org Chem* **21** 413 1956.] Its IR (film) has ν<sub>□□□</sub> at 1480, 1460, 1300, 1184, 1065 and 988-970cm<sup>-1</sup> [Harvey & Mayhood *Can J Chem* **33** 1552 1955].

**Tetramethylsilane (TMS)** [75-76-3] **M 88.2, b 26.3°, d<sub>4</sub><sup>20</sup> 0.639, n<sub>D</sub><sup>20</sup> 1.359.** Distil it from conc H<sub>2</sub>SO<sub>4</sub> (after shaking with it) or LiAlH<sub>4</sub>, through a 5ft vacuum-jacketed column packed with glass helices into an ice-cooled condenser, then percolate it through silica gel to remove traces of halide. [For preparation on a 250g scale see Whitmore & Sommer *J Am Chem Soc* **68** 481 1946, *Beilstein* **4** IV 3875.]

**2,4,6,8-Tetramethyl tetra vinyl cyclotetrasiloxane** [2554-06-5] **M 344.7, m -43.5°, b 111-112°/10mm, 145-146°/13mm, 224-224.5°/758mm, d<sub>4</sub><sup>30</sup> 0.98, n<sub>D</sub><sup>30</sup> 1.434.** A 7ml sample can be distilled in a small Vigreux column at atmospheric pressure without polymerisation or decomposition. It is soluble in cyclohexane. [Kantor et al. *J Am Chem Soc* **77** 1685 1955, *Beilstein* **4** IV 4184.]

**Tetraphenylarsonium chloride hydrate** [507-28-8] **M 418.8, m 258-260°, 261-263°.** A neutralised aqueous solution is evaporated to dryness. The residue is extracted into absolute EtOH, evaporated to a small volume and precipitated by addition of absolute Et<sub>2</sub>O. It is again dissolved in a small volume of absolute EtOH or ethyl acetate and re-precipitated with Et<sub>2</sub>O. *Alternatively,* it is purified by adding conc HCl to precipitate the chloride dihydrate. Redissolve in water, neutralise with Na<sub>2</sub>CO<sub>3</sub> and evaporate to dryness. The residue is extracted with

$\text{CHCl}_3$  and finally crystallised from  $\text{CH}_2\text{Cl}_2$  or EtOH by adding  $\text{Et}_2\text{O}$ . If the aqueous layer is somewhat turbid treat it with Celite and filter it through filter paper. It can be dehydrated before use in a vacuum. The *tetrafluoroborate* salt has *m* 293-295° (needles from MeCN), and the *picrate* salt has *m* 203-204° (from EtOH). [Blicke et al. *J Am Chem Soc* **57** 702 1935, Duke & Brown *J Am Chem Soc* **76** 1443 1954, Popov & Humphrey *J Am Chem Soc* **81** 2043 1959, Singhal & Raj *Synth Inorg Met-org Chem* **23** 1011 1993, *Beilstein* **16** III 1006, **16** IV 1170.] **POISONOUS**.

**Tetraphenylarsonium iodide** [7422-32-4] **M 510.2**. It crystallises from MeOH. [Blicke et al. *J Am Chem Soc* **57** 702 1935, Chatt et al. *J Chem Soc* 1192 1940.] **POISONOUS**.

**Tetraphenylarsonium perchlorate** [3084-10-4] **M 482.8**, **pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. It crystallises from MeOH. [Horner et al. *Chem Ber* **101** 2903 1968.] **POISONOUS and possibly explosive**.

**Tetraphenylbiphosphine (Ph<sub>2</sub>P-PPh<sub>2</sub>)** [1101-41-3] **M 370.4**, **m 120.5° (evacuated tube)**, **120-122° (sealed tube)**, **b 258-260°/1mm**. **pK<sub>Est</sub> >0.0**. This useful precursor of diphenylphosphine compounds is made by heating a mixture of diphenylphosphine (6.0g, 32.5mmol, 829-85-6) and chlorodiphenylphosphine (7.1g, 32.5mmol, 1709-66-9) in petroleum ether (100ml, b 90-100°, freshly distilled over Na) under N<sub>2</sub> at reflux for 3.5 hours during which time all the HCl has evolved. The white crystalline biphosphine that separates (~80%) on cooling is collected, washed with ligroin, dried *in vacuo* and recrystallised or distilled at high vacuum, preferably in a N<sub>2</sub> atmosphere. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, toluene, CCl<sub>4</sub>, and pyridine (yellow solution), and slightly soluble in Et<sub>2</sub>O, EtOH and ligroin. The determined molecular weight was 391.6, 374.9 (method of Rast). When it is heated for 3 hours at 250-300° under N<sub>2</sub>, phosphorus is liberated. The brown residue is boiled for 1 hour with dilute aqueous NaOH and 3% of H<sub>2</sub>O<sub>2</sub>, filtered, cooled, the solid that is collected gives Ph<sub>3</sub>PO (*m* 153° and mixed *m* with authentic Ph<sub>3</sub>PO) after crystallisation from petroleum ether (b 90-100°). *Diphenylphosphonic acid* [*m* 191-192° and mixed melting point with authentic Ph<sub>2</sub>P(O)OH] can be isolated from the mother liquors, making a total yield of products almost quantitative. Reaction of a suspension of the biphosphine (3.3g) in CCl<sub>4</sub> under N<sub>2</sub> with Br (1ml) in CCl<sub>4</sub> is decolourised as it becomes clear. The mixture is then distilled and the oily residue is redistilled to give *bromodiphenylphosphine* (**b 146.5-148°/2.5mm**) as a colourless oil in almost quantitative yield. Bubbling dry air through a suspension of the biphosphine (3.3g) in dry toluene (30ml) at ~0° for 3 hours causes the mixture to become yellow in colour, and after dilution with toluene (30ml) and boiling for a short period it becomes clear. On cooling, and recrystallising the solid that separates from fresh toluene provides the white *dioxide* Ph<sub>2</sub>P(O)-P(O)Ph<sub>2</sub> (**m 167°**, evacuated tube, 1.9g, 53%), with a molecular weight of 418.5, 409.8 (measured by the method of Rast; the required value is 402.4). [Kuchen & Buchwald *Chem Ber* **91** 2871 1958, *Beilstein* **16** H 2871.]

*Sodium dipenylphosphine (Ph<sub>2</sub>PNa)* is readily prepared by adding Na to a solution of the diphosphine in dry Et<sub>2</sub>O or THF (exothermic!), whereby the colour becomes yellow and the orange Na salt separates. It is sensitive to moisture (giving Ph<sub>2</sub>PH) and to CO<sub>2</sub> (giving Ph<sub>2</sub>PCOOH); and the necessary precautions have to be taken. It is a useful reagent for preparing a variety of Ph<sub>2</sub>P-R derivatives and it is best to prepare it freshly when required. [Kuchen & Buchwald *Chem Ber* **92** 227 1959.]

**Tetraphenylboron potassium** see potassium tetraphenylborate.

**Tetraphenylphosphonium chloride** [2001-45-8] **M 374.9**, **m 273-275°**. Crystallise the chloride from acetone and dry at 70° under vacuum. It can also be recrystallised from a mixture of 1:1 or 1:2 dichloromethane/petroleum ether, the solvents having been dried over anhydrous K<sub>2</sub>CO<sub>3</sub>. The purified salt is dried at room temperature under a vacuum for 3 days, and at 170° for a further 3 days. It also crystallises from isoPrOH/Et<sub>2</sub>O or EtOH/Et<sub>2</sub>O. *Extremely hygroscopic*. [Wittig & Geissler *Justus Liebigs Ann Chem* **580** 44, 50 1953, Willard et al. *J Am Chem Soc* **70** 737 1948, *Beilstein* **16** III 851, **16** IV 984.]

**Tetraphenylsilane** [1048-08-4] **M 336.4**, **m 231-233°, 234-235°, 236-238°**. It crystallises from \*benzene as clear colourless bladed needles. It decomposes at ~360°/~760mm on attempted distillation. [George et al. *J Am Chem Soc* **77** 6647 1955, Polis *Chem Ber* **98** 1540 1885, Drew & Landuist *J Chem Soc* 1480 1935, *s Beilstein* **16** H 901, **16** I 525, **16** II 606, **16** III 1199, **16** IV 1372.]

**Tetrasodium pyrene-1,3,6,8-tetrasulfonate** [59572-10-0] M 610.5. Recrystallise this salt from aqueous acetone [Okahata et al. *J Am Chem Soc* **108** 2863 1986].

**Thexyl dimethyl chlorosilane (dimethyl-[2,3-dimethyl-2-butyl] chlorosilane)** [67373-56-2] M 178.8, b 55-56°/10mm, 158-159°/720mm,  $d_4^{20}$  0.970,  $n_D^{20}$  1.428. Purify this chlorosilane by fractional distillation, and store it in small aliquots in sealed ampoules. It is very sensitive to moisture and is estimated by dissolving an aliquot in excess of 0.1M NaOH and titrating with 0.1M HCl using methyl red as indicator [Szabó et al. *Helv Chim Acta* **67** 2128 1984].

**N-(Thexyl dimethylsilyl)dimethylamine (N-[2,3-dimethyl-2-butyl]dimethylsilyl dimethyl-amine)** [81484-86-8] M 187.4, b 156-160°/720mm. Dissolve the amine in hexane, filter, evaporate and distil. It is a colourless oil which is extremely sensitive to moisture. It is best to store small quantities in sealed ampoules after distillation. For estimation of purity, crush an ampoule in excess 0.1N HCl and titrate the excess acid with 0.1M NaOH using methyl red as indicator. [Szabó et al. *Helv Chim Acta* **67** 2128 1984.]

**Tribenzyl chlorosilane** [18740-59-5] M 336.9, m 139-142°, 141-142°, b 300-360°/100mm. It is recrystallised three times from petroleum ether (in slender colourless needles, m 141°). It is sparingly soluble in cold petroleum ether but is soluble in Et<sub>2</sub>O. It does not fume in moist air but is decomposed by H<sub>2</sub>O to give *tribenzyl silanol* m 106° (from petroleum ether). [Robinson & Kipping *J Chem Soc* **93** 439 1908, Jenkins & Post *J Org Chem* **15** 556 1950, *Beilstein* **16** H 906, **16** IV 1498.]

**Tribenzyl phosphine** [76650-89-7] M 304.4, m 96-101°, b 203-210°/0.5mm, pK<sub>Est</sub> ~8.8. Dissolve it in Et<sub>2</sub>O, dry it over Na<sub>2</sub>SO<sub>4</sub>, evaporate and distil it in an inert atmosphere. The distillate solidifies on cooling and is sublimed at 140°/0.001mm. This has m 92-95° (evacuated capillary). When air is bubbled through an Et<sub>2</sub>O solution, it is oxidised to *tribenzyl phosphine oxide*, m 209-212° (evacuated capillary) (it crystallises from Me<sub>2</sub>CO). [Hinton & Mann *J Chem Soc* 2835 1959, *Beilstein* **16** H 771, **16** IV 961.]

**Tri-*n*-butyl borate** [688-74-4] M 230.2, b 110°/11mm, 136°/30mm, 232.4°,  $d_4^{20}$  0.857,  $n_D^{20}$  1.4092. The chief impurities are *n*-butyl alcohol and boric acid (from hydrolysis). It must be handled in a dry-box and can readily be purified by fractional distillation, under reduced pressure. [O'Brien *Aust J Chem* **10** 91 1957, Gerrard & Lappert *J Chem Soc* 2545, 2547 1951, *Beilstein* **1** IV 1544.]

**Tri-*n*-butyl chlorosilane** [995-45-9] M 234.9, b 93-94°/4.5mm, 134-139°/16mm, 250-252°/atm, 142-144°/29mm,  $d_4^{20}$  0.88,  $n_D^{20}$  1.447. Fractionally distil this silane, and store it in small aliquots in sealed ampoules. [Gilman et al. *J Am Chem Soc* **74** 1361 1952, Osthoff & Clark *J Org Chem* **24** 219 1959, *Beilstein* **4** IV 4072.]

**Tri-*n*-butyl phosphate (butyl phosphate)** [126-73-8] M 266.3, m -80°, b 47°/0.45mm, 98°/0.1mm, 121-124°/3mm, 136-137°/5.5mm, 166-167°/17mm, 177-178°/27mm, 289°/760mm (some dec),  $d_4^{20}$  0.980,  $n_D^{20}$  1.44249. The main contaminants in commercial samples are organic pyrophosphates, mono- and di- butyl phosphates and butanol. It is purified by washing successively with 0.2M HNO<sub>3</sub> (three times), 0.2M NaOH (three times) and water (three times), then fractionally distilled under vacuum. [Yoshida *J Inorg Nucl Chem* **24** 1257 1962.] It has also been purified *via* its uranyl nitrate addition compound, obtained by saturating the crude phosphate with uranyl nitrate. This compound is crystallised three times from *n*-hexane by cooling to -40°, and then decomposed by washing with Na<sub>2</sub>CO<sub>3</sub> and water. Hexane is removed by steam distillation; the water is then evaporated under reduced pressure, and the residue is distilled under reduced pressure. [Siddall & Dukes *J Am Chem Soc* **81** 790 1959.]

*Alternatively*, wash it with water, then with 1% NaOH or 5% Na<sub>2</sub>CO<sub>3</sub> for several hours, then finally with water. Dry it under reduced pressure and fractionate it carefully under vacuum. It is a stable colourless oil, sparingly soluble in H<sub>2</sub>O (1ml dissolves in 165ml of H<sub>2</sub>O), but freely miscible in organic solvents. [Kuivila & Masterton *J Am Chem Soc* **74** 4953 1952, Cox & Westheimer *J Am Chem Soc* **80** 5441 1958, <sup>31</sup>P NMR: Van Wazer *J Am Chem Soc* **78** 5715 1956, Fertig et al. *J Chem Soc* 1488 1957, *Beilstein* **1** IV 1531.]

**Tri-*n*-butyl phosphine (TBP)** [998-40-3] M 202.3, b 109-110°/10mm, 115-116°/12mm, 149.5°/50mm, 240.4-

**242.2°/atm,  $d_4^{20}$  0.822,  $n_D^{20}$  1.4463,  $pK_{Est}$  ~7.6.** Fractionally distil TBP under reduced pressure in an inert atmosphere ( $N_2$ ) through an 8 inch gauze-packed column (**b** 110-111°/10mm), and redistil it in a vacuum, then seal it in thin glass ampoules. It is easily oxidised by air to *tri-n-butylphosphine oxide*, **b** 293-296°/745mm. It has a characteristic odour, it is soluble in EtOH, Et<sub>2</sub>O, and \*C<sub>6</sub>H<sub>6</sub>, but is insoluble in H<sub>2</sub>O; and is less easily oxidised by air than the lower molecular weight phosphines. It forms complexes, e.g. with CS<sub>2</sub> (1:1) **m** 65.5° (from EtOH). [Davies & Jones *J Chem Soc* 33 1929, Chernick & Skinner *J Chem Soc* 1401 1956, *Beilstein* 4 IV 3436.] Unlike with Ph<sub>3</sub>P, (MeO)<sub>3</sub>P, (Me<sub>2</sub>N)<sub>3</sub>P or Me<sub>3</sub>P, when (*n*-Bu)<sub>3</sub>P is complexed with DEAD, best results are achieved in the Mitsunobu lactonisation of the final stage in the synthesis of the strained lactone (-)-Echinospurin (XK-213) [Smith et al. *J Am Chem Soc* 114 2567 1992].

**Tri-*tert*-butylphosphine** [13716-12-6] **M 202.3, m ~30-35°, 62.3°, b 103-103°/13mm,  $d_4^{20}$  0.834,  $pK^{25}$  11.4.** It is prepared from *tert*-butylmagnesium chloride and PCl<sub>3</sub> in Et<sub>2</sub>O and finally purified by vacuum fractionation and distillation. [Ger Pat, IG Farbenind DRP 730 638 1939, DRP *Org Chem* 3 1144.] A 1.0M solution in toluene is commercially available (**d<sup>25</sup> 0.861 g/l**). [*Beilstein* 4 III 1771.] It is used with Pd<sub>2</sub>(dba)<sub>3</sub> and Bu<sub>3</sub>SnF in the Pd-catalysed regiocontrolled  $\alpha$ -arylation of trimethylsilyl enol ethers with aryl halides [Iwama & Rawal *Org Lett* 8 5725 2006], is a useful ligand for Pd-catalysed Suzuki-type coupling of arylboronic acids with phenyliodonium ylides of hydroxyquinones [Kazantzi et al. *Synlett* 2597 2006] and is a good ligand for Ni(acac)<sub>2</sub> in catalysing the cross-coupling of aryl- and heteroaryl- halides with aryl Grignard reagents [Böhm et al. *Angew Chem. Int Ed* 39 1602 2000].

**Tri-*tert*-butylphosphine oxide** [6866-70-2] **M 218.3, m 64-69°, 70°, 77°.** The oxide has been prepared by treating *t*-Bu<sub>3</sub>P (1.7g, 85mmol) in Et<sub>2</sub>O (20ml) containing a catalytic amount of KI (0.1g) with 30% of H<sub>2</sub>O<sub>2</sub> until the colour of iodine disappears. The mixture is diluted with H<sub>2</sub>O, basified with aqueous NaOH, the ethereal layer is evaporated, and the phosphine oxide (1.5g, 81%, **m 77°**) is obtained by sublimation at 50°/0.1mm. It is a crystalline hygroscopic solid that is soluble in most organic solvents, and is stable to 250°. Its IR (Nujol) has  $\nu_{max}$  at 1160 (P=O, s) and 815 (P-C, br s) cm<sup>-1</sup>; the <sup>1</sup>H NMR (C<sub>6</sub>D<sub>6</sub>, TMS) has  $\delta_H$  at 0.59 (d, <sup>3</sup>J<sub>P-H</sub> = 12.1Hz, C-Me<sub>3</sub>); the <sup>13</sup>C NMR (C<sub>6</sub>D<sub>6</sub>, TMS) has  $\delta_C$  at 29.08 (s, CH<sub>3</sub>), 38.95 (d, <sup>1</sup>J<sub>P-C</sub> = 51.27, C), and its <sup>31</sup>P NMR (C<sub>6</sub>D<sub>6</sub>, external H<sub>3</sub>PO<sub>4</sub>) has  $\delta$  at 60.86. [Schmidbauer & Blaschke *Z Naturforsch* 33B 1556 1978, Rankin et al. *JCS Dalton Trans* 827 1985.]

**Tri-*n*-butyl phosphite** [102-85-2] **M 250.3, b 114-115°/5mm, 122°/12mm, 130°/17mm, 137°/26mm,  $d_4^{20}$  0.926,  $n_D^{20}$  1.4924.** Fractionate the phosphite through an efficient column. It is stable in air but is slowly hydrolysed by H<sub>2</sub>O. [Gerrard *J Chem Soc* 1464 1940, Fertig et al. *J Chem Soc* 1488 1957, Fields *J Am Chem Soc* 80 2358 1958, Gillis et al. *J Am Chem Soc* 80 2999 1958, *Beilstein* 1 IV 1527.]

**B-Trichloroborazine [B-trichloroborazole]** [933-18-6] **M 183.1, m 83.9-84.5°(sealed evacuated cappillary), 87°, b 88-92°/21mm,  $d_4^{25}$  1.58.** Purify the borazine by distillation from mineral oil. It sublimes at 70°/1mm. [Brown & Laubangayer *J Am Chem Soc* 77 3699 1955, Emelús & Videla *J Chem Soc* 1306 1959.] It is extremely sensitive to moisture and reacts with H<sub>2</sub>O exothermically to give boric acid and NH<sub>4</sub>Cl. Store it in sealed tubes. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, cyclohexane, CS<sub>2</sub>, CHCl<sub>3</sub>, CCl<sub>4</sub>, and C<sub>6</sub>H<sub>5</sub>Cl without decomposition, but in MeOH or EtOH it reacts vigorously, liberating HCl. It is insoluble in pyridine and C<sub>6</sub>H<sub>5</sub>NO<sub>2</sub>. [*Beilstein* 4 III 174, 1 IV 305.]

**Trichloromethyl trimethylsilane (trimethylsilyl trichloromethane)** [5936-98-1] **M 191.6, m 130-132°, b 146-156°/749mm.** This silane distils at atmospheric pressure without decomposition and readily sublimes at 70°/10mm. It has one peak in the <sup>1</sup>H NMR spectrum (CD<sub>2</sub>Cl<sub>2</sub>) with  $\delta$  at 0.38. [Speier *J Am Chem Soc* 73 824 1951, Hergott & Simchen *Synthesis* 626 1980, *Beilstein* 4 IV 3892.]

**Trichlorosilane (SiHCl<sub>3</sub>, silicochloroform)** [10025-78-2] **M 135.4, m -126°, b 31.8°/atm, 32-34°/atm, 36.5°/atm,  $d_4^{25}$  1.342,  $n_D^{20}$  1.4020.** This volatile trichlorosilane can be obtained from silane and HCl in the presence of AlCl<sub>3</sub>, and is purified by fractional distillation at atmospheric pressure in the strict absence of moisture. However, SiHCl<sub>3</sub> is best prepared by placing finely powdered silicon (purified by boiling with HCl

and dilute HF) ground with *ca* 10% of CuCl<sub>2</sub> in a pyrex glass tube (fitted with an adapter and condenser which should extend into the middle of the distilling flask) and heating (carefully at first) in a furnace while a slow stream of absolutely dry HCl gas (from a gas cylinder, or generated from NaCl and concentrated H<sub>2</sub>SO<sub>4</sub>) is passed over the mixture with a furnace temperature of 300°. The receiver is cooled by an acetone/Dry-ice mixture, and the crude product is distilled from the receiver. HCl comes off first then SiHCl<sub>3</sub> distils over at 36.5° in 50% yield. With careful fractionation, the forerun SiH<sub>2</sub>Cl<sub>2</sub>, can be recovered. At 300°, AlCl<sub>3</sub> disproportionates SiHCl<sub>3</sub> to SiH<sub>2</sub>Cl<sub>2</sub> and SiCl<sub>4</sub>. [Note that by using an 4:1 mixture of H<sub>2</sub>/HCl the yield of SiH<sub>2</sub>Cl<sub>2</sub>, [4109-96-0] M 101.0, m -122°, b 8.5°/atm, d<sup>-122°</sup> 1.22, can be greatly improved; and SiH<sub>3</sub>Cl, has m -118°, b -30.5°/atm, d<sup>-113°</sup> 1.15. SiHCl<sub>3</sub> is a water clear, very volatile liquid which fumes in moist air as it is hydrolysed by H<sub>2</sub>O, but is remarkably inert towards metals even the alkali metals such as Na. It is a good reducing agent, is soluble in most organic solvents, e.g. \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub>, and CS<sub>2</sub>, but not in protic solvents in which it may decompose. It is **TOXIC, should not be inhaled**, and should be handled under an efficient fume hood. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 691-692 1963, Fritz *Z Anorg Allg Chem* **280** 134 1955.]

**Tricyclohexylphosphine** [2622-14-2] M 280.4, m 82-83°, pK<sub>Est</sub>~9.5. It recrystallises from EtOH [Boert et al. *J Am Chem Soc* **109** 7781 1987]. [Beilstein **16** IV 947.]

**Triethoxysilane** [998-30-1] M 164.3, m -170°, b 131.2-131.8°/atm, 131.5°/760mm, d<sub>4</sub><sup>20</sup> 0.98753, n<sub>D</sub><sup>20</sup> 1.4377. Fractionate it using a column packed with glass helices of *ca* 15 theoretical plates in an inert atmosphere. Store it in aliquots in sealed ampoules because it is sensitive to moisture. [Spauschus et al. *J Am Chem Soc* **72** 1377 1950, MacKenzie et al. *J Am Chem Soc* **72** 2032 1950, Havill et al. *J Org Chem* **13** 280 1948, Beilstein **1** IV 1359.]

**Triethylborane** [97-94-9] M 146.0, m -92.5°, b 94-97°, 94-95°, n<sub>D</sub><sup>20</sup> 1.378, d<sub>4</sub><sup>20</sup> 0.678. This borane distils at 56-57°/220mm. It can also be purified *via* its *ammonia addition complex* which is distilled in a high vacuum, decomposed with dry HCl, and the Et<sub>3</sub>B is distilled out. It is commercially available as a 15% solution in hexane or as 1M solution in hexane. [Brown *J Am Chem Soc* **67** 376 1945, Bamford et al. *J Chem Soc* 471 1946, Lin et al. *J Organomet Chem* **312** 277 1986, Beilstein **4** III 1957, **4** IV 4359.]

**Triethyl borate** [150-46-9] M 146.0, b 44.5°/45mm, 118°, n<sub>D</sub><sup>20</sup> 1.378, d<sub>4</sub><sup>20</sup> 0.864. Dry the ester over sodium, then distil it. Also fractionate it through a gauze packed column. [Charnlet et al. *J Chem Soc* 2288 1952, as for tributyl borate Johnson & Tompkins *Org Synth Coll Vol II* 106 1943, Beilstein **1** III 1339, **1** IV 1365.]

**Triethyl phosphate** [78-40-0] M 182.2, b 40-42°/0.25-0.3mm, 98-98.5°/8-10mm, 90°/10mm, 130°/55mm, 204°/680mm, 215-216°/760mm, d<sub>4</sub><sup>25</sup> 1.608, n<sub>D</sub><sup>20</sup> 1.4053. Dry the phosphate by refluxing it with solid BaO and then fractionally distil it under reduced pressure. It is kept over Na and distilled. Store it in the receiver protected from light and moisture. *Alternatively*, it is dried over Na<sub>2</sub>SO<sub>4</sub> and distilled under reduced pressure. The middle fraction is stirred for several weeks over anhydrous Na<sub>2</sub>SO<sub>4</sub> and again fractionated under reduced pressure until the specific conductance reaches a constant low value of κ<sup>25</sup> 1.19 x 10<sup>8</sup>, κ<sup>40</sup> 1.68 x 10<sup>8</sup>, and κ<sup>55</sup> 2.89 x 10<sup>8</sup> ohm<sup>-1</sup> cm<sup>-1</sup>. It has also been fractionated carefully under reduced pressure through a glass helices-packed column. It is soluble in EtOH, Et<sub>2</sub>O and H<sub>2</sub>O (dec). [Estok & Wendlandt *J Am Chem Soc* **77** 4767 1955, Hoffmann et al. *J Am Chem Soc* **78** 6413 1956. (P NMR), Muller et al. *J Am Chem Soc* **78** 3557 1956, French et al. *J Chem Soc* 3582 1959, IR: Bellamy & Beecher *J Chem Soc* 475 1952 and McIvor *Can J Chem* **36** 820 1958, Kosolapoff *Organophosphorus Compounds*, Wiley p 258 1950, Beilstein **1** IV 1339.]

**Triethylphosphine** [554-70-1] M 118.2, b 100°/7mm, 127-128°/744mm, d<sub>4</sub><sup>15</sup> 0.812, n<sub>D</sub><sup>18</sup> 1.457, pK<sup>25</sup> 8.69 (also available as a 1.0M solution in THF). All operations should be carried out in an efficient fume cupboard because it is flammable, toxic and has a foul odour. Purify the phosphine by fractional distillation at atmospheric pressure in a stream of dry N<sub>2</sub>, as it is oxidised by air to the oxide. In 300% excess of CS<sub>2</sub> it forms Et<sub>3</sub>PCS<sub>2</sub> (m 118-120° crystallising from MeOH) which decomposes in CCl<sub>4</sub> to give Et<sub>3</sub>PS as a white solid m 94° when recrystallised from EtOH. [Sorettas & Isbell *J Org Chem* **27** 273 1962, Henderson & Streuli *J Am*

*Chem Soc* **82** 5791 1960, pK: Henderson & Streuli *J Am Chem Soc* **82** 5791 1960, see also trimethylphosphine.] Store it in sealed vials under N<sub>2</sub>.

Alternatively, dissolve it in Et<sub>2</sub>O and shake it with a solution of AgI and KI to form the insoluble complex. Filter off the complex, dry it over P<sub>2</sub>O<sub>5</sub> and the Et<sub>3</sub>P is regenerated by heating the silver iodide complex in a tube attached to a vacuum system. It has the odour of hyacinths. [Hewitt & Holliday *J Chem Soc* 530 1953, Schettas & Isbell *J Org Chem* **27** 2573 1962, Kosolapoff *Organophosphorus Compounds*, Wiley p 31 1950, *Beilstein* **4** IV 3431.]

**Triethyl phosphite** [122-52-1] **M 166.2, b 48-49°/11mm, b 52°/12mm, 57.5°/19mm, 157.9°/757mm, d<sub>4</sub><sup>20</sup> 0.9687, n<sub>D</sub><sup>20</sup> 1.4135.** Treat the ester with Na (to remove water and any dialkyl phosphonate), then decant and distil it under reduced pressure, with protection against moisture; or distil it in a vacuum through an efficient Vigreux column or a column packed with Penn State 0.16 x 0.16inch protruded nickel packing and a variable volume take-off head. [Ford-Moore & Perry *Org Synth Coll Vol IV* 955 1963, Kosolapoff *Organophosphorus Compounds*, Wiley p 203 1950, *Beilstein* **1** IV 1333.]

**Triethyl phosphonoacetate (triethyl carboxymethyl phosphonate)** [867-13-0] **M 224.2, b 83-84°/0.5mm, 103°/1.2mm, 143-144°/11mm, 260-262°/atm, d<sub>4</sub><sup>20</sup> 1.1215, n<sub>D</sub><sup>20</sup> 1.4310.** Purify the phosphono-acetate by fractional distillation, preferably *in vacuo*. The <sup>31</sup>P NMR has a P resonance at 19.5 relative to orthophosphate. [Kosolapoff & Powell *J Am Chem Soc* **68** 1103 1946, Kosolapoff & Powell *J Am Chem Soc* **72** 4198 1950, Speziale & Freeman *J Org Chem* **23** 1586 1958, *Beilstein* **4** IV 3613.]

**Triethyl phosphonofornate** [1474-78-8] **M 210.2, b 70-72°/0.1mm, 122.5-123°/8mm, 130-131°/10mm, 138.2°/12.5mm, d<sub>4</sub><sup>20</sup> 1.22, n<sub>D</sub><sup>20</sup> 1.423.** Dissolve it in Et<sub>2</sub>O, shake this with H<sub>2</sub>O (to remove any trace of NaCl impurity), dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and distil it using an efficient fractionating column. [Nylén *Chem Ber* **57** 1035 1924, Reetz et al. *J Am Chem Soc* **77** 3813 1955, Monson *Advanced Organic Synthesis* Academic Press p 89 1972, *Beilstein* **3** II 103.]

**Triethyl 2-phosphonopropionate [ethyl 2-(diethoxyphosphinyl)propionate]** [3699-66-9] **M 238.2, b 76-77°/0.2mm, 137-138.5°/17mm, d<sub>0</sub><sup>20</sup> 1.096, n<sub>D</sub><sup>20</sup> 1.432.** Purify the ester by fractional distillation with high reflux ratio, preferably using a spinning band column. [Kosolapoff & Powell *J Am Chem Soc* **72** 4198 1950, Kresze et al. *Justus Liebigs Ann Chem* **756** 112 1972, *Beilstein* **4** IV 3617.]

**Triethylsilane** [617-86-7] **M 116.3, b 105-107°, 107-108°, d<sub>0</sub><sup>20</sup> 0.734, n<sub>D</sub><sup>20</sup> 1.414.** Reflux triethylsilane over molecular sieves, then distil it. It is passed through neutral alumina before use [Randolph & Wrighton *J Am Chem Soc* **108** 3366 1986]. [*Beilstein* **4** IV 3895.]

**Triethylsilyl-1,4-pentadiene (1,4-pentadien-3-yloxy-trimethylsilane)** [62418-65-9] **M 198.4, b 72-74°/12mm, d<sub>4</sub><sup>20</sup> 0.842, n<sub>D</sub><sup>20</sup> 1.439.** Dissolve the diene in pentane, wash this with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate, and distil it under vacuum. R<sub>F</sub> values on Kieselgel 60 are 0.15 (pentane) and 0.60 (\*C<sub>6</sub>H<sub>6</sub>). [IR, NMR, MS: Oppolzer et al. *Helv Chim Acta* **64** 2002 1981.]

**Tri-*n*-hexylborane** [1188-92-7] **M 265.3, b 127°/1.5mm, b 185-188°/30mm.** Treat the borane with hex-1-ene and 10% anhydrous Et<sub>2</sub>O for 6 hours at gentle reflux under N<sub>2</sub>, then distil it in a vacuum through an 18 inch glass helices-packed column under N<sub>2</sub> taking the fraction **b** 130°/2.1mm to 137°/1.5mm. The distillate may still contain some di-*n*-hexylborane [Brown & Subba Rao *J Am Chem Soc* **81** 6423 1959, Mirviss *J Am Chem Soc* **83** 3051 1961]. [*Beilstein* **4** IV 4362.]

**Triisomyl phosphate** [919-62-0] **M 308.4, b 143°/3mm.** Purify the ester by repeated crystallisation of its addition compound with uranyl nitrate from hexane. Decompose the complex, and distil the ester at high vacuum. [Siddall *J Am Chem Soc* **81** 4176 1959.] [see *tributyl phosphate* and Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol **6** pp 211-577 1973.]

**Triisobutyl phosphate** [126-71-6] **M 266.3, b 119-129°/8-12mm, 192°/760mm, d<sub>4</sub><sup>20</sup> 0.962, n<sub>D</sub><sup>20</sup> 1.421.** Purify the phosphate by repeated crystallisation of its addition compound with uranyl nitrate from hexane. (see *tributyl*

phosphate.) [Siddall *J Am Chem Soc* **81** 4176 1959; see Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol **6** pp 211-577 1973.]

**Triisooctyl thiophosphate** [30108-39-5] **M 450.6**. Purify the ester by passing its solution in  $\text{CCl}_4$  through a column of activated alumina. [See Ailman & Magean in *Organo Phosphorus Compound* (Kosolapoff & Maier eds) Wiley Vol **7** pp 487-465 1973, and tri-*p*-tolyl phosphate below.]

**Tri-isopropyl borate** [(iso-PrO)<sub>3</sub>B, boric acid tri-isopropyl ester, Boron iso-propoxide] [5419-55-6] **M 188.1**, **b** 75°/76mm, 90°/120mm, 139-141°/760mm,  $d_4^{25}$  0.815,  $n_D^{20}$  1.376. The borate ester is prepared in 85% yield from *iso*-propanol (250ml) and  $\text{NaBH}_4$  (6.84g) followed by dropwise addition of AcOH (0.18g) over a period of 11 minutes, and then refluxing for 4 hours (fume cupboard and a "Dry-Ice" condenser as 16.15L of  $\text{H}_2$  are released, CARE due to its flammability). Then fractionate the ester through a Widmer column. [Brown et al. *J Am Chem Soc* **78** 3613 1956.] The ester is a good reagent for borylation, e.g. *ortho*-borylating 1-substituted naphthalenes, by reacting with 1-halo- or cyano- naphthalenes (after treatment with LiTMP) to form 2-boryl esters which are later used in Pd-catalysed cross-coupling reactions [Lysén et al. *Synthesis* 3478 2006]. [Beilstein **1** H 363, **1** II 382, **1** III 1468, **1** IV 1488.]

**Triisopropyl phosphite** [116-17-6] **M 208.2**, **b** 58-59°/7mm, 63-64°/7mm,  $n_D^{25}$  1.4082. Distil the phosphite from sodium, under vacuum, through a column packed with glass helices. (This removes any dialkyl phosphonate.) [Ford-Moore & Williams *J Chem Soc* 1465 1947, Arbuzov *Chem Ber* **38** 1171 1905, see Verkade & Coskren in *Organo Phosphorus Compound* (Kosolapoff & Maier eds) Wiley Vol **2** pp 1-187 1972, Beilstein **1** IV 1476.]

**Trimesitylphosphine** [23897-15-6] **M 388.5**, **m** 205-206°,  $pK_{\text{Est}} \sim 8.0$ . It recrystallises from EtOH [Boert et al. *J Am Chem Soc* **109** 7781 1987]. The *P*-methyl iodide has **m** 269° (yellow powder from EtOH or  $\text{H}_2\text{O}$ ). [Beilstein **16** H 774.]

**Trimethyllyl phosphate** [14019-81-9] **M 260.3**, **b** 134.5-140°/5mm,  $n^{25}$  1.4454. Purify it as for triisoamyl phosphate. [Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol **6** pp 211-577 1973.]

**Trimethoxysilane** [2487-90-3] **M 122.2**, **m** -114.8°, 81.1°/760mm, 84°/atm,  $d_4^{20}$  0.957,  $n_D^{20}$  1.359. Likely impurities are  $\text{Si}(\text{OMe})_4$  and  $\text{H}_2\text{Si}(\text{OMe})_2$ . Efficient fractionation is essential for removing these impurities [IR: Sternbach & MacDiarmid *J Am Chem Soc* **81** 5109 1959, Heilfrich & Hausen *Chem Ber* **57** 795 1924, Beilstein **1** IV 1266.]

**Trimethyl borate (methylborate, trimethoxyboron)** [121-43-7] **M 103.9**, **m** -34°, **b** 67-68°/742mm,  $d_4^{20}$  0.928,  $n_D^{20}$  1.3610. Carefully fractionate the borate through a gauze-packed column. Re-distil and collect it in weighed glass vials and seal them. Keep it away from moisture. It undergoes alkyl exchange with alcohols and forms azeotropes, e.g. with MeOH the azeotrope consists of 70%  $(\text{MeO})_3\text{B}$  and 30% MeOH with **b** 52-54°/760mm,  $d$  0.87. [Charnley et al. *J Chem Soc* 2288 1952, Gerrard & Lappert *Chem Ind (London)* 53 1952, Schlesinger et al. *J Am Chem Soc* **75** 213 1953.] It has also been dried with Na and then distilled. [Beilstein **1** IV 1269.]

**Trimethyl boroxine ( $\text{Me}_3\text{B}_3\text{O}_3$ )** [823-96-1] **M 125.5**, **m** -38°, **b** 80°/742mm, 79.3°/755mm,  $d_4^{20}$  0.902,  $n_D^{20}$  1.362. Possible impurity is methylboronic acid. If present, then add a few drops of conc  $\text{H}_2\text{SO}_4$  and distil it immediately, then fractionate it through an efficient column. [McCusker et al. *J Am Chem Soc* **79** 5179 1957, IR: Goubeau & Keller *Z Anorg Allgem Chem* **272** 303 1953, Beilstein **4** IV 4378.]

**Trimethylphenylsilane (phenyltrimethylsilane)** [768-32-1] **M 150.3**, **b** 67.3°/20mm, 98-99°/80mm, 170.6°/738mm,  $d_4^{25}$  0.8646,  $n_D^{20}$  1.491. Fractionally distil the silane at atmospheric or reduced pressure (Podbielniak column) and estimate it by GC with a column packed with Silicone Fluid No 710 on Chromosorb P support. [Gilman et al. *J Org Chem* **18** 1743 1953, Maienthal et al. *J Am Chem Soc* **76** 6392 1954, House & Respass *J Organomet Chem* **4** 95 1965, Roberts et al. *J Am Chem Soc* **71** 2923 1949, Freiser et al. *J Am Chem*

*Soc* 75 2821 1953, *Beilstein* 16 I 525, 16 II 605, 16 III 1198, 16 IV 1361.]

**Trimethyl phosphate** [512-56-1] **M 140.1**, **b** 77°/12mm, 94°/22mm, 110°/60mm, 197.2°/atm,  $d_4^{20}$  1.0213,  $n_D^{20}$  1.3961. Purify the phosphate by fractionation through an efficient column at high reflux ratio. It is quite soluble in H<sub>2</sub>O; the solubility is 1:1 at 25°. [Becker *J Am Chem Soc* 74 2923 1952, IR: Bergmann et al. *J Chem Soc* 847 1952, McIvor et al. *Can J Chem* 36 820 1958, Kosolapoff *Organophosphorus Compounds*, Wiley p 258 1950, and Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973, *Beilstein* 1 IV 1259.]

**Trimethylphosphine** [594-09-2] **M 76.1**, **m** -86°, **b** 38-39°/atm, **pK<sup>25</sup> 8.65**, (also available as a 1.0M solution in THF or toluene). **All operations should be carried out in an efficient fume cupboard because it is flammable, toxic and has a foul odour.** Distil trimethylphosphine at atmospheric pressure in a stream of dry N<sub>2</sub> (apparatus should be held together with springs to avoid loss of gas from increased pressure in the system) and the distillate run into a solution of AgI in aqueous KI whereby the silver complex [Me<sub>3</sub>PAGI]<sub>4</sub> separates steadily. Filter off the complex, wash it with saturated aqueous KI solution, then H<sub>2</sub>O and dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. The dry complex is heated in a flask (in a stream of dry N<sub>2</sub>) in an oil bath at 140°, when pure Me<sub>3</sub>P distils off (bath temperature can be raised up to 260°). The vapour pressure of Me<sub>3</sub>P at 20° is 466mm and the **b** is 37.8° [Thomas & Eriks *Inorg Synth* IX 59 1967]. Alternatively, freshly distilled Me<sub>3</sub>P (6g) is shaken with a solution of AgI (13.2g, 1.1mol) in saturated aqueous KI solution (50ml) for 2 hours. A white solid, not wetted with H<sub>2</sub>O, separates rapidly. It is collected, washed with the KI solution, H<sub>2</sub>O, and dried [Mann et al. *J Chem Soc* 1829 1937]. The silver complex is stable if kept dry in the dark, in which state it can be kept indefinitely. Me<sub>3</sub>P can be generated from the complex when required. Store it under N<sub>2</sub> in a sealed container. It has been distilled in a vacuum line at -78° *in vacuo* and condensed at -96° [IR and NMR: Crosbie & Sheldrick *J Inorg Nucl Chem* 31 3684 1969]. The pK<sup>22</sup> by NMR was 8.80 [Silver & Lutz *J Am Chem Soc* 83 786 1961, pK<sup>25</sup> 8.65: Henderson & Strueuli *J Am Chem Soc* 82 5791 1960].

The [Me<sub>2</sub>PAGI]<sub>4</sub> complex [12389-34-3] is a flammable solid which has **m** 140-142°. It is decomposed by heating gently in one arm of an inverted U tube. The other arm is kept in a freezing mixture. The complex dissociates, and pure Me<sub>3</sub>P collects in the cold arm and is used at once. It should not be allowed to come in contact with air [for AsMe<sub>3</sub> see Mann & Wells *J Chem Soc* 708 1938]. The CS<sub>2</sub> complex has **m** 119° (crystallising from 95% EtOH) and decomposes in CCl<sub>4</sub> to give Me<sub>3</sub>PS **m** 154° (from EtOH) [Soretas & Isbell *J Org Chem* 27 273 1962].

*Trimethylphosphine hydrochloride* is unstable and volatilises at 75°/0.4mm (120°/14mm). [Brown *J Am Chem Soc* 67 503 1945, IR: Wagstaffe & Thompson *Trans Faraday Soc* 40 41 1944, Kosolapoff *Organo-phosphorus Compounds*, Wiley p 31 1950, *Beilstein* 4 IV 3429.]

**Trimethyl phosphite** [121-45-9] **M 124.1**, **b** 22°/23mm, 86-86.5°/351mm, 111-112°/760mm, 111°/atm,  $d_4^{20}$  1.0495,  $n_D^{20}$  1.408. Treat the phosphite with Na (to remove water and any dialkyl phosphonate), then decant and distil it with protection against moisture. It has also been treated with sodium wire for 24 hours, then distilled in an inert atmosphere onto activated molecular sieves [Connor et al. *J Chem Soc, Dalton Trans* 511 1986]. It can be fractionally distilled using a spinning band column at high reflux ratio. It is a colourless liquid which is slowly hydrolysed by H<sub>2</sub>O. [Gillis et al. *J Am Chem Soc* 80 2999 1958, <sup>31</sup>P NMR: Callis et al. *J Am Chem Soc* 79 2719 1957, Kosolapoff *Organophosphorus Compounds*, Wiley p 203 1950, *Beilstein* 1 IV 1256.]

**Trimethylsilyl acetamide** [13435-12-6] **M 131.3**, **m** 38-43°, 52-54°, **b** 84°/13mm, 185-186°/atm. Distil the amide repeatedly in an inert atmosphere with all operations to be performed in an anhydrous atmosphere. In the presence of moisture, trimethylsilanol (**b** 31-34°/26mm) is formed and is a likely impurity (check by NMR). [Birkofer et al. *Chem Ber* 96 1473 1963, *Beilstein* 4 IV 4011.]

**Trimethylsilyl acetonitrile (TMSAN)** [18293-53-3] **M 113.2**, **b** 49-51°/10mm, 65-70°/20mm,  $d_4^{20}$  0.8729,  $n_D^{20}$  1.4420. Check if NMR and IR spectra show impurities; if present dissolve it in \*C<sub>6</sub>H<sub>6</sub> (10 volumes), wash it with buffer (AcOH/AcONa pH ca 7) several times, dry (CaCl<sub>2</sub>) it, filter, evaporate and distil it. Its IR (CCl<sub>4</sub>) has

$\nu_{\max}$  at 2215 (CN)  $\text{cm}^{-1}$ , and  $^1\text{H}$  NMR ( $\text{CCl}_4$ ) with  $\delta$  at 0.23 (s, 9H,  $\text{SiMe}_3$ ), and 1.53 (s, 2H,  $\text{CH}_2\text{CN}$ ). [Matsuda et al. *J Chem Soc, Perkin Trans 1* 26 1979, *Beilstein* 4 IV 3974.]

**Trimethylsilyl azide** [4648-54-8] **M 115.2, b 92-95°/atm, 95-99°/atm,  $d_4^{20}$  0.878,  $n_D^{20}$  1.441.** Distil the azide through a Vigreux column in a  $\text{N}_2$  atmosphere maintaining the oil bath temperature thermostat at 135-140°. Check the purity by  $^1\text{H}$  NMR [ $\text{CHCl}_3$ ,  $\delta$ : single peak at 13cps from  $\text{Me}_4\text{Si}$ ]. Likely impurities are siloxane hydrolysis products. The azide is thermally stable even at 200° when it decomposes slowly without explosive violence. All the same, it is advisable to carry out the distillation behind a thick safety screen in a fumehood because unforeseen **EXPLOSIVE** azides may be formed on long standing. [Birkofer & Wagner *Org Synth Coll Vol VI* 1030 1988.]

**Trimethylsilyl chloride (TMCS trimethyl chlorosilane, chlorotrimethylsilane)** [75-77-4] **M 108.6, b 56-57°/atm, 58°/760mm,  $d_4^{20}$  0.86,  $n_D^{20}$  1.388.** Likely impurities are other chlorinated methylsilanes and tetrachlorosilane (b 57.6°), some of which can form azeotropes. To avoid the latter, very efficient fractional distillation is required. It has been fractionated through a 12 plate glass helices-packed column with only the heart-cut material being used. It has also been fractionated through a 90cm, 19mm diameter Stedman column. Purify it by redistilling from  $\text{CaH}_2$  before use. [Sauer et al. *J Am Chem Soc* 70, 4254 1948, Sauer & Hadsell *J Am Chem Soc* 70 4258 1948, Langer et al. *J Org Chem* 23 50 1958, *Beilstein* 4 IV 4007.] **FLAMMABLE and CORROSIVE.**

**Trimethylsilyl chloroacetate** [18293-71-5] **M 166.7, m -20°, b 57-58°/14mm, 70-71°/30mm, 159°/760mm,  $d_4^{20}$  1.057,  $n_D^{20}$  1.4231.** Purify the chloroacetate by repeated fractionation, and taking the fractions with clean NMR spectra. [Anderson *J Am Chem Soc* 74 2371 1952, *Beilstein* 4 IV 4004.]

**Trimethylsilyl cyanide** [7677-24-9] **M 99.2, m 8-11°, 10.5-11.5°, 11-12°, 12-12.5°, b 54-55°/87mm, 67-71°/168mm, 114-117°/760mm, 118-119°/760mm,  $d_4^{20}$  0.79  $n_D^{20}$  1.43916.** The material should have only one sharp signal in the  $^1\text{H}$  NMR (in  $\text{CCl}_4$  with  $\text{CHCl}_3$  as internal standard) with  $\delta$  at 0.4, and the IR with  $\nu_{\max}$  at  $2210\text{cm}^{-1}$  (C≡) [McBride & Beachall *J Am Chem Soc* 74 5247 1952, Prober *J Am Chem Soc* 77 3224 1955]; otherwise purify it by fractionating through an 18 x 1/4inch column. [Evers et al. *J Am Chem Soc* 81 4493 1959.] It has also been carefully distilled using a 60cm vacuum jacketed column. If the volume of sample is small, the cyanide can be chased (in the distillation) with xylene that had been previously distilled over  $\text{P}_2\text{O}_5$ . It is **HIGHLY TOXIC** and **FLAMMABLE**. [Evans et al. *J Org Chem* 39 914 1974, *Beilstein* 4 IV 3893.]

**Trimethylsilyldiazomethane (diazomethyl)trimethylsilane** [18107-18-1] **M 114.2, distils between 0°/100mm and 40°/15mm.** This diazo-silane is non-explosive, and non-mutagenic [Aoyama & Shiori *Chem Pharm Bull Jpn* 29 3249 1981] and the careful precautions used in preparing diazomethane are not strictly necessary here. It is a good safe and stable substitute for diazomethane and excess of reagent can be evaporated off. It is available commercially as 2.0M solutions in  $\text{Et}_2\text{O}$  or hexane. Solutions in hexane can be stored for periods of more than 6 months at 0° in the absence of light without noticeable decomposition. Care should be taken, however, when opening a container, as it may become pressurised by  $\text{N}_2$  which may be released, and should be done at ~0°.

Several methods were used for preparing this diazo-silane and some start from trimethylsilylmethylmagnesium chloride, e.g. by reaction with diazomethane (7-74%, Lappert et al. *J Chem Soc (A)* 2954 1970, Martin *Synth Commun* 13 809 1983), or with  $\text{TsN}_3$  (~17%, Mori et al. *Chem Pharm Bull Jpn* 30 3380 1982, Barton & Hoekman *React Inorg. Met-Org Chem* 9 297 1979], but the transfer of the azo group is most practical, high yielding and used on a large scale when diphenyl phosphorazidate is used, and is described here.

Trimethylsilylmethylmagnesium chloride [13170-43-9] is prepared in an inert atmosphere (e.g. argon) under strictly anhydrous conditions from Mg turning (10.7g, 0.44g-atom), anhydrous  $\text{Et}_2\text{O}$  (40ml) and 1,2-dibromoethane (0.1ml to initiate reaction) which are stirred at ~25° for 15 minutes, then 10ml of a mixture of trimethylsilylmethyl chloride (45.4g, 370mmol, freshly distilled at 97°/atm, see [2344-80-1] below) in anhydrous  $\text{Et}_2\text{O}$  (100ml) is added with stirring all at once to start the reaction followed by the rest in a dropwise manner at such a rate as to maintain gently reflux during the addition (~2 hours). After the exothermic reaction subsides, reflux and stirring are continued for 1 hour then the mixture is cooled to ~25°, and this Grignard

reagent is used in the next step. Under strictly anhydrous conditions and in a maintained argon atmosphere, a solution of diphenyl phosphorazidate (91.2g, 330mmol, freshly distilled at 134-136°/0.2mm see [26386-88-9] above) in anhydrous Et<sub>2</sub>O (350ml) is cooled and stirred in an ice-NaCl bath until the temperature is -10°, and the preceding Grignard reagent is added dropwise (*via* a funnel or a cannula) at such a rate as to maintain the inner temperature of the solution below 0° (addition requires ~ 1.5 hours, forming a large amount of white precipitate when two-thirds of the reagent is added). After addition is complete, the ice-salt bath is replaced by an ice bath while stirring for 2 hours, then stirring is continued for 14-16 hours. The mixture is again cooled in an ice-salt bath until the inner temperature drops to -15°, and cold H<sub>2</sub>O (35ml) is added dropwise while keeping the inner temperature below 0° (requiring ~ 1 hour), and the resulting yellow silyldiazomethane mixture is stirred for a further 0.5 hours. The mixture is filtered, the white solid is washed with Et<sub>2</sub>O (3 x 100ml), the combined filtrate and washings are washed with cold H<sub>2</sub>O (2 x 100ml), dried (Na<sub>2</sub>SO<sub>4</sub>), and the filtrate is placed in a 1L flask equipped with a Teflon-coated magnetic stirrer bar and a 30cm Vigreux column (1.5cm diameter). The mixture is slowly concentrated to *ca* 200ml by distillation at ~760mm with a bath temperature below 45° during ~6 hours (the colour of the distilled Et<sub>2</sub>O is yellow due to some co-distillation of the diazo-silane. This rate of distillation is important as a faster rate results in decreased yields of the diazo-silane. The concentration time can be reduced to ~ 4 hours if a 30cm Widmer column is used. The remaining deep-yellow solution is then distilled, in the same equipment, by reducing the pressure to 100mm at 0° (bath temperature) and then at 15mm and 40° (bath temperature) while the distillate is collected in a receiver in a Dry-ice/Me<sub>2</sub>CO bath until distillation is complete. The distillate is dried (MgSO<sub>4</sub>), filtered and diluted with redistilled hexane (100ml). This solution is then concentrated again by distillation through the 30cm Vigreux column until the temperature of the vapour reaches 68° (with final oil bath temperature at 87°) which requires ~3 hours. About 80-110ml of the residual yellow hexane solution contains 220-230mmol of *trimethylsilyldiazomethane* (67-70% based on the phosphorazidate). Its IR (hexane) has  $\nu_{\max}$  at 2075, 1260 and 885 cm<sup>-1</sup>, the <sup>1</sup>H NMR (100MHz, hexane, CHCl<sub>3</sub> as internal standard) has  $\delta$  at 0.16 (s, 9H, -Si(CH<sub>3</sub>)<sub>3</sub>) and 2.58 (s, 1H, -CHN<sub>2</sub>), and the <sup>1</sup>H NMR (100MHz, \*C<sub>6</sub>H<sub>6</sub>) has  $\delta$  at -0.03 (s, 9H, -Si(CH<sub>3</sub>)<sub>3</sub>) and 2.23 (s, 1H, -CHN<sub>2</sub>).

The concentration of *trimethylsilyldiazo-methane* in hexane can be determined by adding 91mg (0.5mmol) of dibenzyl (see [103-29-7], Aromatic compounds) in 1ml of the hexane solution containing it, and measuring the <sup>1</sup>H NMR. The concentration of the azo-silane (*x* mmol/ml) is calculated from the formula:  $x = 2a/b$ , where *a* is the integral value (mm) of the methine proton ( $\delta$ : 2.58) of *trimethylsilyldiazomethane*, and *b* is the integral value (mm) of the methylene protons ( $\delta$ : 2.99) of dibenzyl. [Shioiri & Aoyama *Org Synth Coll Vol VIII* 612 1993, Mori et al. *Chem Pharm Bull Jpn* **30** 3380 1982].

It has been used for stereo and/or regio selective dipolar cycloadditions, e.g. with chiral acrylamides it produces <sup>2</sup> $\Delta$ -pyrazoline carboxylic acids (azaprolines) after protio-desilylation [Mish et al. *J Am Chem Soc* **119** 8379 1997], or with diethyl *trans*-glutaconate where the pyrazoline ester is oxidised to the respective pyrazole [Di & Rein *Tetrahedron Lett* **45** 4703 2004]. It also acts as a carbene source in stereoselective [1 + 4] annulation reactions with silylvinylketenes to form a variety of cyclopent-2-enones in good to excellent yields [Moser et al. *J Org Chem* **71** 6542 2006].

**2-Trimethylsilyl-1,3-dithiane** [13411-42-2] **M 192.2, b 54.5°/0.17mm, 100°/8mm,  $d_4^{20}$  1.04,  $n_D^{20}$  1.533.** Fractionally distil the dithiane through an efficient column and collect the fractions that have the correct NMR and IR spectra. Its <sup>1</sup>H NMR (CCl<sub>4</sub>) has  $\tau$  at 6.36 (SiMe<sub>3</sub>), 9.87 (SCHS) and dithiane H at 7 and 8 (ratio 1:9:4:2) from Me<sub>4</sub>Si; the UV has  $\lambda_{\max}$  at 244nm ( $\epsilon$  711), sh 227nm ( $\epsilon$  800). [Corey et al. *J Am Chem Soc* **89** 434 1967.]

**2-(Trimethylsilyl)ethanesulfonyl chloride (SES-Cl)** [106018-85-3] **M 200.8, b 60°/0.1mm, 146.8°/760mm,  $d_4^{25}$  1.059,  $n_D^{25}$  1.4444.** Check IR: if the bands at ~3200 (OH) cm<sup>-1</sup> are strong, then much of the SES-Cl had hydrolysed, and it should be treated with POCl<sub>3</sub> (with cooling) and stirred at ~25° for about 1 hour, poured into ice cold H<sub>2</sub>O, extracted with CH<sub>2</sub>Cl<sub>2</sub>, washed with NaHCO<sub>3</sub>, dried (Na<sub>2</sub>SO<sub>4</sub>), evaporated, and it distils as a yellow oil in a vacuum. This procedure is used for converting the Na salt [18143-30-1] to SES-Cl. It reacts with amines to form amides, e.g. SES-NRR', which on heating with CsF (i.e. F<sup>-</sup> ions) in DMF at 95° provide the amine (NHRR'), SO<sub>2</sub> and CH<sub>2</sub>=CH<sub>2</sub> [Weinreb et al. *Tetrahedron Lett* **27** 2099 1986]. [Ribiere et al. *Chem Rev*

106 2249 2006.]

**Trimethylsilyl ethanol** [2916-68-9] **M 118.3, b 53-55°/11mm, 75°/41mm, 95°/100mm,  $d_4^{25}$  0.8254,  $n_D^{25}$  1.4220.** If the NMR spectrum is not clean, then dissolve the alcohol in Et<sub>2</sub>O, wash it with aqueous NH<sub>4</sub>Cl solution, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and distil it. The 3,4-dinitrobenzoyl derivative has **m 66°** (from EtOH). [NMR: Speier et al. *J Am Chem Soc* **79** 974 1957, *Z Naturforsch* **14b** 137 1959, *Beilstein* **4** IV 3951.]

**2-(Trimethylsilyl)ethoxymethyl chloride (SEMCl)** [76513-69-4] **M 166.7, b 57-59°/8mm,  $d_4^{20}$  0.942,  $n_D^{20}$  1.4350.** Dissolve SEMCl in pentane, dry it (MgSO<sub>4</sub>), evaporate and distil the residual oil in a vacuum. Stabilise it with 10 ppm of diisopropylamine. Store it under N<sub>2</sub> in a sealed container in a refrigerator. [Lipshutz & Pegram *Tetrahedron Lett* **21** 3343 1980.]

**2-(Trimethylsilyl)ethoxymethyltriphenylphosphonium chloride** [82495-75-8] **M 429.0, m 140-142°, 145-149°.** Wash the solid with AcOH and recrystallise it from CH<sub>2</sub>Cl<sub>2</sub>/EtOAc. Dry it in a vacuum desiccator. *Hygroscopic.* The <sup>1</sup>H NMR (CDCl<sub>3</sub>) has δ at -0.2 (s, Me<sub>3</sub>Si), 0.8 (t, 8Hz, CH<sub>2</sub>Si), 3.83 (t, 8Hz, OCH<sub>2</sub>), 5.77 (d, *J*<sub>PH</sub> = 4Hz, P<sup>+</sup>-CH<sub>2</sub>O) and 7.70 (m, aromatic H). [Schönauer & Zbiral *Justus Liebigs Ann Chem* 1039 1983.]

**Trimethylsilylethyl phenylsulfone (phenyl-2-trimethylsilylethylsulfone)** [73476-18-3] **M 242.4, m 52°.** Dissolve the sulfone in Et<sub>2</sub>O, wash it with saturated HCO<sub>3</sub> followed by saturated NaCl, H<sub>2</sub>O and dried (MgSO<sub>4</sub>). Filtration followed by evaporation leaves residual crystals with **m 52°**. [Hsiao & Shechter *Tetrahedron Lett* **23** 1963 1982, Bortolini et al. *J Org Chem* **53** 2688 1985.]

**Trimethylsilyl isocyanate** [1118-02-1] **M 115.2, b 90-92°/atm, b 91.3-91.6°/atm,  $d_4^{20}$  0.850  $n_D^{20}$  1.43943.** Purify it by repeated fractionation as for the isothiocyanate below. [Eaborn *J Chem Soc* 3077 1950, *Beilstein* **4** III 1861, **4** IV 4011.]

**Trimethylsilyl isothiocyanate** [2290-65-5] **M 131.3, m -33°, b 142.6-143.1°/759mm, 143.8°/760mm,  $n_D^{20}$  1.4809.** The <sup>1</sup>H NMR spectrum should have only one peak; if not, purify it by repeated fractionation in an all-glass system using a 50cm (4mm internal diameter) column without packing. [Anderson *J Am Chem Soc* **69** 3049 1947, Fehér & Blümcke *Chem Ber* **90** 1934 1957, Neidleim & Hege *Synthesis* 51 1975, *Beilstein* **4** III 1861, **4** IV 4011.]

**(Trimethylsilyl)methanol** [3219-63-4] **M 104.2, b 120-122°/754mm, 122-123°/768mm,  $d_4^{20}$  0.83  $n_D^{20}$  1.4176.** If the NMR indicates impurities (should have only two signals), then dissolve it in Et<sub>2</sub>O, shake this with aqueous 5N NaOH, M H<sub>2</sub>SO<sub>4</sub>, saturated aqueous NaCl, dry (MgSO<sub>4</sub>) and distil it using an efficient column at atmospheric pressure. The 3,5-dinitrobenzoate has **m 70-70.5°** (from 95% EtOH). [Huang & Wang *Acta Chem Sin* **23** 291 1957, cf. *Chem Abstr* **52** 19911 1958, Speier et al. *J Am Chem Soc* **81** 1844 1959 and Speier et al. *J Am Chem Soc* **70** 1117 1948, *Beilstein* **4** III 1844, **4** IV 2876.]

**(Trimethylsilyl)methylamine (aminomethyl trimethylsilane)** [18166-02-4] **M 103.2, b 101.6°/735mm,  $d_4^{20}$  0.77,  $n_D^{20}$  1.416.** A possible contaminant is hexamethyldisiloxane. It should have two <sup>1</sup>H NMR signals in CDCl<sub>3</sub>; if not, dissolve it in \*C<sub>6</sub>H<sub>6</sub>, shake it with 15% aqueous KOH, separate, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and distil it using a still of *ca* 10 theoretical plates. The water azeotrope has **b 83°/735mm**; hence it is important to dry the extract well. The *hydrochloride* has **m 198-199°** (from MeOH or Me<sub>2</sub>CO). [Noll et al. *J Am Chem Soc* **73** 3867 1951, *Beilstein* **4** IV 3878.]

**(Trimethylsilyl)methyl chloride (chloromethyltrimethylsilane, silico-neopentyl chloride)** [2344-80-1] **M 122.7, b 97.1°/734mm, 98.2-98.7°/747mm  $d_4^{25}$  0.979,  $n_D^{20}$  1.4180.** This chloride is prepared by stirring a mixture of TMS (174g, 2 moles, Me<sub>4</sub>Si see [75-76-3]) in dry CCl<sub>4</sub> (150ml) and PCl<sub>5</sub> (3g) [use suitably placed traps of Dry-ice/Me<sub>2</sub>CO to prevent loss of TMS] under a reflux condenser and irradiated with a 450 watt GE sunlight lamp while dry chlorine is bubbled through for 4 hours at a rate of 0.5 mole per hour (check weight of flask occasionally). The mixture is then fractionated using a 15-plate glass helix packed column to give the *trimethylsilylmethyl chloride* (53g, 0.44 mole), **b 97.1°/734mm**, *polychlorinated tetramethylsilane* (93g) and

recovered TMS (62g, 0.7 mole). The *chloride* can be analysed by placing ~0.2g in a gelatin capsule in a Parr bomb and fused with Na<sub>2</sub>O<sub>2</sub> (15g) and sucrose (1g). The melt is treated with H<sub>2</sub>O, acidified with HNO<sub>3</sub> and the chloride ion is determined by titration using the Volhard method. [Whitmore & Sommer *J Am Chem Soc* **68** 481 1946, *Beilstein* **4** IV 3877.] It has also been prepared by reaction of chloro-dimethyl-chloromethyl-silane and MeMgBr in Et<sub>2</sub>O [Whitmore et al. *J Am Chem Soc* **69** 1976 1947, Roedel *J Am Chem Soc* **71** 271 1949]. [*Beilstein* **4** III 1844, **4** IV 3877.]

It is readily converted to *trimethylsilylmethylmagnesium chloride* (see [13170-43-9] above) in ~90% yield. It reacts with *n*-BuLi in pentane to give *trimethylsilylmethylLi* [Sommer et al. *J Am Chem Soc* **71** 2750 1949], which reacts with anhydrous cerium (III) chloride to form *trimethylsilylmethyl cerium dichloride*. The latter reacts with acyl halides to provide bis-β-silylethyl tertiary alcohols that efficiently undergo a trimethylchlorosilane-promoted Peterson reaction to generate allylsilanes in high overall yields [Anderson & Fuchs *Synth Commun* **17** 621 1987]. It is useful also for Peterson olefination and homologation of aldehydes and ketones *via* 1,2-epoxysilanes [Lee et al. *Tetrahedron* **45** 5877 1989]. *Trimethylsilylmethyl bromide* [18243-41-9] has **M** 167.1, **b** 115.5°/atm, **d**<sub>4</sub><sup>25</sup> 1.17, **n**<sub>D</sub><sup>20</sup> 1.444 [*Beilstein* **4** IV 3878], and *trimethylsilylmethyl iodide* [4206-67-1] has **M** 214.1, **b** 139-141°/atm, **d**<sub>4</sub><sup>25</sup> 1.433, **n**<sub>D</sub><sup>20</sup> 1.491 [*Beilstein* **4** IV 3878], and both should be stored in the dark.

**Trimethylsilylmethyl phenylsulfone (phenyltrimethylsilylmethylsulfone)** [17872-92-3] **M** 228.4, **m** 28-32°, **b** 121°/0.01mm, 160°/6mm, **n**<sub>D</sub><sup>20</sup> 1.5250. Fractionate the sulfone at high vacuum and recrystallise it from pentane at -80°. If too impure (*cf* IR), dissolve it in CH<sub>2</sub>Cl<sub>2</sub> (*ca* 800ml for 100g), wash this with 2M aqueous NaOH (2 x 200ml), brine, dry, evaporate and distil it. [Craig et al. *J Chem Soc, Perkin Trans 1* 1949 1985, IR and NMR: Cooper *J Am Chem Soc* **76** 3713 1954.]

**1-Trimethylsilyloxy-1,3-butadiene** [6651-43-0] **M** 142.3, **b** 131°/760mm (mixture of isomers), 49.5°/25mm (*E*-isomer), **d**<sub>4</sub><sup>20</sup> 0.8237, **n**<sub>D</sub><sup>20</sup> 1.447. Purify the butadiene by fractional distillation, and collect the fractions with the required <sup>1</sup>H NMR. Store it under N<sub>2</sub> — it is a flammable and moisture-sensitive liquid. [Caseau et al. *Bull Soc Chim Fr* 16658 1972, Belge Patent 670,769, *Chem Abstr* **65** 5487d 1966.]

**1-(Trimethylsilyloxy)cyclopentene** [19980-43-9] **M** 156.3, **b** 45°/11mm, 75-80°/20-21mm, **d**<sub>4</sub><sup>20</sup> 0.878, **n**<sub>D</sub><sup>20</sup> 1.441. If too impure as seen by the NMR spectrum, then dissolve it in 10 volumes of pentane, shake with cold NaHCO<sub>3</sub> (3 x 500ml), then 1.5M HCl (200ml) and aqueous NaHCO<sub>3</sub> (200ml) again, dry (Na<sub>2</sub>SO<sub>4</sub>), filter, evaporate and distil it through a short Vigreux column. Its <sup>1</sup>H NMR (CDCl<sub>3</sub>) has δ at 0.21 (s, 9H), 1.55 (m, 2H), 1.69 (m, 2H), 2.05 (br d, 4H) and 4.88 (br s, 1H). GLPC in a 6ft x 1/8inch with 3% SP2100 on 100-120 mesh Supelcoport column should give one peak. Store dry. [For the cyclohexene analogue see Varghese et al. *Org Synth Coll Vol VIII* 460 1993.]

**2-(Trimethylsilyloxy)furan** [61550-02-5] **M** 156.3, **b** 34-35°/9-10mm, 42-50°/17mm, 40-42°/25mm, **d**<sub>4</sub><sup>20</sup> 0.950, **n**<sub>D</sub><sup>20</sup> 1.436. Fractionally distil the furan using a short path column. Its <sup>1</sup>H NMR in CCl<sub>4</sub> has δ at 4.90 (dd, *J* = 1.3Hz, 3H), 6.00 (t, *J* = 3Hz, 4H) and 6.60 (m, 5H). [Yoshii et al. *Heterocycles* **4** 1663 1976.]

**4-Trimethylsilyloxy-3-penten-2-one (cis) (acetylacetone enol trimethylsilyl ether)** [13257-81-3] **M** 172.3, **b** 66-68°/4mm, 61-63°/5mm, **d**<sub>4</sub><sup>20</sup> 0.917, **n**<sub>D</sub><sup>20</sup> 1.452. Fractionally distil the enone, and store it in glass ampoules which are sealed under N<sub>2</sub>. It hydrolyses readily in contact with moisture giving, as likely impurities, hexamethyldisiloxane and 2,4-pentanedione. [West *J Am Chem Soc* **80** 3246 1958, *Beilstein* **4** IV 4003.]

**1-(Trimethylsilyl)-2-phenylacetylene (1-phenyl-2-trimethylsilylacetylene)** [78905-09-6] **M** 174.3, **b**\* 45-46°/0.1mm, 67°/5mm, 87.5°/9mm, **d**<sub>4</sub><sup>20</sup> 0.8961, **n**<sub>D</sub><sup>20</sup> 1.5284. Dissolve the acetylene in Et<sub>2</sub>O, wash with H<sub>2</sub>O, dry and fractionate it through a Todd column. [Benkeser & Hicker *J Am Chem Soc* **80** 5298 1958.]

**3-(Trimethylsilyl)propyne** [13361-64-3] **M** 112.3, **b** 99-100°/760mm, **d**<sub>4</sub><sup>20</sup> 0.7581, **n**<sub>D</sub><sup>20</sup> 1.4091. Fractionally distil the propyne, and add 2,6-di-*tert*-butyl-*p*-cresol (~0.5%) to stabilise it. [Petrov et al. *Doklady Acad Nauk USSR* **93** 293 1953, *cf Chem Abstr* **48** 13616 1954, *Beilstein* **4** IV 3938.]

**2-Trimethylsilylpyridine** [13737-04-7] **M** 151.3, **b** 47-49°/5mm, 69°/14.5mm, 74°/21mm, **d**<sub>4</sub><sup>25</sup> 0.9113, **n**<sub>D</sub><sup>20</sup>

**1.489.** Purify it by distillation in a vacuum, but if it is discoloured then dissolve it in Et<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>, wash it with H<sub>2</sub>O, dry over Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil the residue. It is more readily hydrolysed than the 3- or 4-trimethylsilyl isomers and the relative rates of hydrolysis in H<sub>2</sub>O:MeOH:EtOH are 740:120:1; pyridine being liberated. Acids inhibit the reaction but bases have little effect. [Anderson et al. *J Chem Soc B* 450 1958, Effenberger & Häbich *Justus Liebigs Ann Chem* 512 149-151 1934, Itami *Tetrahedron* 57 5045 2001.]

**3-Trimethylsilylpyridine** [13779-37-2] **M151.3, b 94°/30mm, d<sub>4</sub><sup>25</sup> 0.9113, n<sub>D</sub><sup>20</sup> 1.4913.** Purify it as for the 1-isomer above except that less care needs to be taken as it hydrolyses more slowly. [Anderson et al. *J Chem Soc B* 450 1958.]

**4-Trimethylsilylpyridine** [18301-46-7] **M151.3, b 107°/48mm, d<sub>4</sub><sup>25</sup> 0.9113, n<sub>D</sub><sup>20</sup> 1.4868.** Purify it as for the 1-isomer above except that less care needs to be taken as it hydrolyses more slowly. [Anderson et al. *J Chem Soc B* 450 1958.]

**1-Trimethylsilyl-1,2,4-triazole** [18293-54-4] **M 141.3, b 74°/12mm, d<sub>4</sub><sup>20</sup> 0.99, n<sub>D</sub><sup>20</sup> 1.4604.** Fractionally distil it at atmospheric pressure in an inert atmosphere because it is moisture sensitive. [Birkofer et al. *Chem Ber* 93 2804 1960.]

**Trimethylsilyl trifluoromethane (trifluoromethyl trimethylsilane, Ruppert's reagent)** [81290-20-2] **M 142.2, b 54-55°, 55-55.5°, d<sub>4</sub><sup>20</sup> 0.962, n<sub>D</sub><sup>20</sup> 1.332.** Purify the silane by distilling it from trap to trap in a vacuum of 20mm using a bath at 45° and Dry-ice/Me<sub>2</sub>CO bath for the trap. The liquid in the trap is then washed with ice cold H<sub>2</sub>O (3x), the top layer is collected, dried (Na<sub>2</sub>SO<sub>4</sub>), and the liquid is decanted and fractionated through a helices-packed column at atmospheric pressure. <sup>1</sup>H, <sup>13</sup>C, <sup>19</sup>F, and <sup>29</sup>Si NMR can be used for assessing the purity of fractions. [Ruppert et al. *Tetrahedron Lett* 25 2195 1984, Krishnamurti et al. *J Org Chem* 56 984 199, Beilstein 4 IV 3892.]

**2,8,9-Trimethyl-2,5,8,9-tetraaza-1-phosphabicyclo[3.3.3]undecane (Verkade Superbase, 2,8,9-trimethyl-pro-azaphosphatrane)** [120666-13-9] **M 216.3, m 110-115°, pKa 26.8 (DMSO).** The free superbase can be prepared in three ways. When tris(dimethylamino)phosphine (8.8g, 54mmol, toxic see [1608-26-0]) and tris[2-(methylamino)-ethyl]amine (10.0g, 53mmol, see [65604-89-9]) are dissolved in dry xylene and refluxed for 21 days, then the solvent is removed *in vacuo*, a residual thick oil is obtained. On heating this oil at 105°/0.05mm the *pro-azaphosphatrane* (53g, 24.5mmol, 46%) sublimes out as a colourless waxy solid.

In a second preparation the *trimethyl-pro-azaphosphatrane hydrochloride* (0.87g, 3.4mmol, see [138800-17-6] below) dissolved in MeCN (10ml) is added dropwise to a suspension of freshly sublimed *tert*-BuOK (0.41g, 3.7mmol) in MeCN (20ml), stirred for 30 minutes, the solvent is removed *in vacuo*, the residue is extracted with hexane (2 x 30ml), evaporated, and the residue is sublimed at 60°/0.01mm to give the pure *superbase* (80%) as a white solid. [Schmidt et al. *Z Anorg Allg Chem* 578 75 1989.]

In a third preparation the *trimethyl-azaphosphatrane hydrochloride* is heated slowly with a large excess of anhydrous NaOH under vacuum. No reaction occurs below 200°, but above this temperature sudden sublimation of the free superbase occurs. Extraction of the sublimate, and/or the reaction mixture, with \*C<sub>6</sub>H<sub>6</sub> followed by evaporation of the solvent provided pure free base (53% yield). Sometimes charring occurs during this reaction resulting in lower yields. [Lensink et al. *J Am Chem Soc* 111 3478 1959.]

The *free base* is an unusually strong Lewis base. It deprotonates phenol (pKa 10), protonated "Proton Sponge" [1,8-(bismethylamino)naphthalene HI (pKa 12.3, Alder et al. *J C S, Chem Commun* 723 1968), diethyl malonate (pKa 13, Pearson & Dillon *J Am Chem Soc* 75 2439 1953), H<sub>2</sub>O (pKa 13.99, Harned & Robinson *Trans Farad Soc* 36 2747 1940; or ? pKa 15.74 Reeve et al. *Can J Chem* 57 2747 1979), but does not react with *tert*-BuOH (pKa 16.5, Reeve et al. *Can J Chem* 57 2747 1979), to give the protonated base as the only product. In the presence of 1 equivalent of H<sub>2</sub>O an equilibrium mixture of free base to cation of ca 4:1 is obtained (from NMR studies) [Lensink et al. *J Am Chem Soc* 111 3478 1989.] The un-methylated analogous free base is not very stable although its hydrochloride is quite stable and it is a *stronger base* than the 2,8,9-trimethyl- or the respective 2,8,9-tribenzyl- derivative. The upper limit of the pKa of the trimethyl compound in DMSO is ~26.8 (by <sup>13</sup>P NMR) [Laramay & Verkade *J Am Chem Soc* 112 9421 1990]. The *free base* has IR (Nujol) with ν<sub>max</sub> at 1332 s, 1303 m, 1244 s, 1226 s, 1197m, 1145 s, 1128 s, 1053 s, 1004 s, 960 w, 887m, 850s, 767w, 650s and 634 s

cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CD<sub>3</sub>CN, TMS) has δ at 2.60 (d, 9H, <sup>3</sup>J<sub>PH</sub> = 11.0 Hz, CH<sub>3</sub>) and 2.76 (br s, 12H, CH<sub>2</sub>); the <sup>13</sup>C NMR (75MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 37.2 (<sup>2</sup>J<sub>PC</sub> = 41.0 Hz, CH<sub>3</sub>), 49.4 (d, <sup>2</sup>J<sub>PC</sub> = 6.7 Hz, N<sub>eq</sub>CH<sub>2</sub>) and 51.3 (s, N<sub>ax</sub>CH<sub>2</sub>); the <sup>31</sup>P NMR (122MHz, C<sub>6</sub>D<sub>6</sub>, 85% H<sub>3</sub>PO<sub>4</sub>) has δ at 120.8; and the HRMS has *m/z* 216.15088 (calc for M is 216.15039). [Schmidt et al. *Z Anorg Allg Chem* **578** 75 1989, Lensink et al. *J Am Chem Soc* **111** 3478 1989, Laramay & Verkade *J Am Chem Soc* **112** 9421 1990, Laramay & Verkade *Z Anorg Allg Chem* **608** 163 1991, Schmidt et al. *Inorg Chem* **29** 2214 1990]. In addition to being easily protonated (see following entry), the free base has been readily converted to P-substituted adducts (*five-coordinated phosphatranes*), where the P-H is replaced by P-X, where -X is -O (by using Me<sub>3</sub>SiO<sub>2</sub>), -S (by using S<sub>8</sub>), -Se (by using Se), -N<sub>3</sub>P (by using PhN<sub>3</sub>), -NPh (by using PhN<sub>3</sub> and heating), -CS<sub>2</sub> (by using CS<sub>2</sub>) and -PtCl<sub>2</sub> [by using (Et<sub>2</sub>S)<sub>2</sub>-PtCl<sub>2</sub>] in 60 to >80% yields [Schmidt et al. *Z Anorg Allg Chem* **578** 75 1989]. **Note** that the difference between the *pro-phosphatranes* and the *phosphatranes* is that the latter possess an apical N<sup>+</sup>O<sup>-</sup> bond as evidenced by the shorter bond distance in X-Ray diffraction analyses and the large decrease in the δ <sup>31</sup>P chemical shifts in the NMR spectra.

**2,8,9-Trimethyl-2,5,8,9-tetraaza-1-phosphabicyclo[3.3.3]undecane hydrochloride (Verkade Superbase HCl, 2,8,9-trimethylazaphosphatrane HCl)** [138800-17-6] **M 252.7, m 132-136°**. Like the preceding free base it is prepared in different ways — two are described here. A solution of PCl<sub>3</sub> (1.22g, 8.85mmol) in CH<sub>2</sub>Cl<sub>2</sub> (5.0ml) is added to a solution of P(NMe) (2.89g, 17.7mmol, toxic see [1608-26-0]) in CH<sub>2</sub>Cl<sub>2</sub> (5.0ml), then cooled to 5°, and TREN (3.88g, 26.5mmol, see [4097-89-6]) in CH<sub>2</sub>Cl<sub>2</sub> (25ml) is added during ~15 minutes. The precipitate formed is filtered off, washed with CH<sub>2</sub>Cl<sub>2</sub> (25ml) and dried to give the pure *hydrochloride* (5.58g, ~100%). Crystals for X-Ray diffraction can be grown from concentrated MeOH solutions. The *2,8,9-tribenzyl hydrochloride* can be similarly prepared in 89% yield by using tris-(*N*-benzylaminoethyl)amine in place of TREN. [Laramay & Verkade *Z Anorg Allg Chem* **605** 163 1991].

In the second preparation tris[2-(methylamino)-ethyl]amine (1.67g, 11.4mmol), see [65604-89-9]) in CH<sub>2</sub>Cl<sub>2</sub> (20ml) is added with stirring to a mixture of ClP(NMe<sub>2</sub>)<sub>2</sub> (1.76g, 11.4mmol, TOXIC) and Et<sub>3</sub>N (1.5g, 15mmol) in CH<sub>2</sub>Cl<sub>2</sub> (30ml) during 5 minutes, then the mixture is stirred at ~25° for 1 hour, the volatiles are evaporated off *in vacuo* to give the *hydrochloride* quantitatively. It can be recrystallised from hexane/CHCl<sub>3</sub> (82% yield) in colourless crystals. When the hydrochloride salt is treated with AgBF<sub>4</sub> in CH<sub>2</sub>Cl<sub>2</sub> the *HBF<sub>4</sub> salt* is obtained quantitatively. The salts are very difficult to deprotonate (see preceding entry). The X-ray crystal structure shows a short intra-annular bond between the apical N atom and the P atom with a robust P—H bond, the positive charge being on the apical N atom. [Laramay & Verkade *J Am Chem Soc* **112** 9421 1990, Lensink et al. *J Am Chem Soc* **111** 3478 1989.] It has <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, TMS) with δ at 2.61 (d, 9H, <sup>3</sup>J<sub>PH</sub> = 17.4 Hz, CH<sub>3</sub>) and 3.03 (dt, 6H, N<sub>ax</sub>CH<sub>2</sub>, <sup>3</sup>J<sub>PH</sub> = 11.0 Hz, <sup>3</sup>J<sub>HH</sub> = 6.2 Hz), 3.58 (dt, 6H, N<sub>eq</sub>CH<sub>2</sub>, <sup>3</sup>J<sub>PH</sub> = 4.7 Hz, <sup>3</sup>J<sub>HH</sub> = 6.2 Hz), 5.20 (d, 1H, <sup>1</sup>J<sub>PH</sub> = 6.2 Hz); the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>) has δ at 34.4 (d, <sup>2</sup>J<sub>PC</sub> = 17.1 Hz, CH<sub>3</sub>), 41.3 (d, <sup>2</sup>J<sub>PC</sub> = 6.1 Hz, N<sub>ax</sub>CH<sub>2</sub>) and 47.3 (d, <sup>2</sup>J<sub>PC</sub> = 7.3 Hz, N<sub>eq</sub>CH<sub>2</sub>); the <sup>31</sup>P NMR (122MHz, CDCl<sub>3</sub>, 85% H<sub>3</sub>PO<sub>4</sub>) has δ at -10.6 [Lensink et al. *J Am Chem Soc* **111** 3478 1989].

**Trimethyl vinyl silane** [754-05-2] **M 100.2, b 54.4°/744mm, 55.5°/767mm, d<sub>4</sub><sup>25</sup> 0.6865, n<sub>D</sub><sup>25</sup> 1.3880**. If the <sup>1</sup>H NMR spectrum shows impurities, then dissolve it in Et<sub>2</sub>O, wash it with aqueous NH<sub>4</sub>Cl solution, dry over CaCl<sub>2</sub>, filter, evaporate and distil it at atmospheric pressure in an inert atmosphere. It is used as a co-polymer and may polymerise in the presence of free radicals. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>. [Nagel & Post *J Org Chem* **17** 1379 1952, Beilstein **4** IV 3922.]

**Tri(4-nitrophenyl)phosphate** [3871-20-3] **M 461.3, m 155-156°, 156°, 156-158°, 157-159°**. This phosphate has been recrystallised from AcOH, dioxane, AcOEt and Me<sub>2</sub>CO and dried it in a vacuum over P<sub>2</sub>O<sub>5</sub>. [Katelaar & Gersmann *J Am Chem Soc* **72** 5777 1950, Moffatt & Khorana *J Am Chem Soc* **79** 3741 1957.]

**Tri(*n*-octyl)phosphine oxide** [78-50-2] **M 386.7, m 59.5-60°, pK<sub>Est</sub> <0**. Mason, McCarty and Peppard [*J Inorg Nuclear Chem* **24** 967 1962] purified the oxide by stirring a 0.1M solution in \*benzene with an equal volume of 6M HCl at 40° in a sealed flask for 48 hours, then washed the \*benzene solution successively with water (twice), 5% aqueous Na<sub>2</sub>CO<sub>3</sub> (three times) and water (six times). The \*benzene and water were then evaporated under reduced pressure at room temperature. Zingaro and White [*J Inorg Nucl Chem* **12** 315 1960]

treated a petroleum ether solution of the oxide with aqueous  $\text{KMnO}_4$  (to oxidise any phosphinous acids to phosphinic acids), then with sodium oxalate,  $\text{H}_2\text{SO}_4$  and  $\text{HCl}$  (to remove any manganese compounds). The petroleum ether solution was slurried with activated alumina (to remove phosphinic acids), filtered, evaporated and the residue was recrystallised from petroleum ether or cyclohexane at  $-20^\circ$ . It can also be recrystallised from  $\text{EtOH}$ . [*Beilstein* 4 IV 3466.]

**Tri(neo-pentyl) phosphate** [14540-59-1] **M 320.4**. Crystallise it from hexane. [See Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973.]

**Triphenylarsine** [603-32-7] **M 306.2, m 60-62°**. Recrystallise  $\text{Ph}_3\text{As}$  from  $\text{EtOH}$  or aqueous  $\text{EtOH}$  [Dahlinger et al. *J Chem Soc, Dalton Trans* 2145 1986, Boert et al. *J Am Chem Soc* 109 7781 1987]. [*Beilstein* 16 H 829, 16 I 431, 16 II 407, 16 III 921, 16 IV 1139.] **HIGHLY TOXIC**.

**Triphenylarsine oxide** [1153-05-5] **M 322.2, m 194-196° (anhydrous)**. The anhydrous oxide crystallises from  $^*\text{C}_6\text{H}_6$ , whereas the *monohydrate* separates from  $\text{EtOH}$  with **m 118° (m 114-117° was also reported)**. [*Beilstein* 16 H 846, 16 I 433, 16 II 433, 16 III 1022, 16 IV 1139.] **HIGHLY TOXIC**.

**Triphenyl borane (borane triphenyl, triphenyl boron)** [960-71-4] **M 242.1, m 134-140°, 137°, 139-141°, 142-142.5°, 147.5-148°, 151°, b 203°/15mm**. Recrystallise the borane three times from  $\text{Et}_2\text{O}$  or  $^*\text{C}_6\text{H}_6$  under  $\text{N}_2$  and dry it at  $130^\circ$ . It can be distilled in a high vacuum at  $300-350^\circ$  and has been distilled (**b 195-215°/-15mm**) in vacuum using a bath temperature of  $240-330^\circ$ .  $\text{N}_2$  is introduced into the apparatus before dismantling. It forms complexes with amines. [Nielsen et al. *Chem Ind (London)* 1069 1957, Wittig et al. *Justus Liebigs Ann Chem* 563 110 1949, Bent & Dorfman *J Am Chem Soc* 57 1259 1935, *Beilstein* 16 IV 1623.]

**Triphenyl phosphate** [115-86-6] **M 326.3, m 49.5-50°, b 245°/11mm**. Crystallise the phosphate from  $\text{EtOH}$  or petroleum ether (b 60-80°)/ $\text{EtOH}$ . [Cox & Westheimer *J Am Chem Soc* 80 5441 1958, Krishnakumar & Sharma *Synthesis* 558 1983, Cherbuliez in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973, *Beilstein* 6 III 658, 6 IV 720.] It has been used as an effective ligand with phase transfer catalysts [Nakoji et al. *J Org Chem* 67 7418 2002].

**Triphenylphosphine** [603-35-0] **M 262.3, m 77-78°, 79°, 79-81°, 80.5°, 80-81°, b >360°(377°, in inert gas),  $d_4^{25}$  1.194,  $d_4^{80}$  1.075 (liq),  $\text{pK}^{25}$  2.73**. The phosphine crystallises from hexane,  $\text{MeOH}$ , diethyl ether,  $\text{CH}_2\text{Cl}_2$ /hexane or 95%  $\text{EtOH}$ . Dry it at  $65^\circ$ / $<1\text{mm}$  over  $\text{CaSO}_4$  or  $\text{P}_2\text{O}_5$ . Chromatograph it through alumina using (4:1)  $^*\text{benzene}/\text{CHCl}_3$  as eluent. [Blau & Espenson et al. *J Am Chem Soc* 108 1962 1986, Buchanan et al. *J Am Chem Soc* 108 1537 1986, Randolph & Wrighton *J Am Chem Soc* 108 3366 1986, Asali et al. *J Am Chem Soc* 109 5386 1987.] It has also been crystallised twice from petroleum ether and 5 times from  $\text{Et}_2\text{O}/\text{EtOH}$  to give crystals with **m 80.5°**. Alternatively, dissolve it in conc  $\text{HCl}$ , and upon dilution with  $\text{H}_2\text{O}$  it separates because it is weakly basic, it is then crystallised from  $\text{EtOH}/\text{Et}_2\text{O}$ . It recrystallises unchanged from  $\text{AcOH}$ . [Forward et al. *J Chem Soc Suppl.* p121 1949, Muller et al. *J Am Chem Soc* 78 3557 1956.]  $3\text{Ph}_3\text{P}\cdot 4\text{HCl}$  crystallises out when  $\text{HCl}$  gas is bubbled through an  $\text{Et}_2\text{O}$  solution, it has **m 70-73°**, but recrystallises very slowly and is deliquescent.  $\text{Ph}_3\text{P}\cdot\text{HBr}$  [6399-81-1] has **m 196° (dec)**. The *hydriodide*, made by adding  $\text{Ph}_3\text{P}$  to hydriodic acid, is not hygroscopic and decomposes at  $\sim 100^\circ$ . The *chlorate* (1:1) salt has **m 165-167°**, but decomposes slowly at  $100^\circ$ . All salts hydrolyse in  $\text{H}_2\text{O}$  to give  $\text{Ph}_3\text{P}$  [IR, UV: Sheldon & Tyree *J Am Chem Soc* 80 2117 1958,  $\text{pK}$ : Henderson & Streuli *J Am Chem Soc* 82 5791 1960, Kosolapoff, *Organophosphorus Compounds*, Wiley 1950]. [*Beilstein* 16 IV 951.]

§ It is also available commercially on a support of polystyrene cross-linked with 2% divinylbenzene as the free base [39319-11-4], loading of  $\sim 3.2$  nmol/g, 100-200 mesh], and as the *hydrobromide* [loading 2.5-3.0 mmol Br, 200-400 mesh].

**Triphenylphosphine dibromide** [1034-39-5] **M 422.1, m 235°, 245-255°(dec)**. The dibromide recrystallises from  $\text{MeCN}/\text{Et}_2\text{O}$ . Although it has been recrystallised from  $\text{EtOH}$ , this is not recommended as it converts alcohols to alkyl bromides as it has brominating power. It deteriorates on keeping, and it is best to prepare it afresh. [Anderson & Frenor *J Am Chem Soc* 86 5037 1964, Horner et al. *Justus Liebigs Ann Chem* 626 26 1959, *Beilstein* 16 III 864.]

**Triphenylphosphine oxide** [791-28-6] **M 278.3, m 152.0°, 156-157°, pK<sub>Est</sub> ~2.10 (aqueous H<sub>2</sub>SO<sub>4</sub>), pK<sup>25</sup> 2.9 (in MeNO<sub>2</sub>).** It crystallises from absolute EtOH and is dried *in vacuo*. The *gold chloride complex* has **m 177.5-178.5°**. [Addison & Sheldon *J Chem Soc* 2705 1956, Cox & Westheimer *J Am Chem Soc* **80** 5441 1958, Beilstein **16** III 864, **16** 1011.]

**Triphenyl phosphite** [101-02-0] **M 310.3, m 16-20°, 21-23°, b 181-189°/1mm, 183-184°/1mm, d<sub>4</sub><sup>20</sup> 1.183.** Its ethereal solution is washed successively with aqueous 5% NaOH, distilled water and saturated aqueous NaCl, then dried with Na<sub>2</sub>SO<sub>4</sub> and distilled under vacuum after evaporating the diethyl ether. [Walsh *J Am Chem Soc* **81** 3023 1959, Verkade & Coskren in *Organo Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973, Beilstein **6** IV 695.] It has been used as an effective ligand with phase transfer catalysts [Nakoji et al. *J Org Chem* **67** 7418 2002].

**Triphenylphosphorylidene acetaldehyde (formylmethylenetriphenylphosphorane)** [2136-75-6] **M 304.3, m 185-187°, 186-187°(dec).** Recrystallise it from Me<sub>2</sub>CO, or dissolve it in \*C<sub>6</sub>H<sub>6</sub>, wash with N NaOH, dry (MgSO<sub>4</sub>), evaporate, and crystallise the residue from Me<sub>2</sub>CO. It can be prepared from its precursor, formylmethyltriphenylphosphonium chloride (which crystallises from CHCl<sub>3</sub>/EtOAc), by treatment with Et<sub>3</sub>N and extraction with \*C<sub>6</sub>H<sub>6</sub>. [Tripett & Walker *J Chem Soc* 1266 1961.]

**(Triphenylphosphoranylidene)ketene [Bestmann ylide (triphenylphosphoranylidene)ethenone, ketenylidene(triphenyl)phosphorane]** [15596-07-3] **M 302.3, m 167-173°, 173°.** This versatile reagent is prepared in a dry inert atmosphere (N<sub>2</sub> or argon) by first mixing NaNH<sub>2</sub> (19.5g, 0.5mole) in toluene (1.3L, dried by passage through alumina) and bis(trimethylsilyl)amine (81g, 105ml, 0.5mol, toxic and corrosive), and heated (70-80° bath temperature) under a reflux condenser until a clear and colourless solution of sodium hexamethyldisilazane (see hexamethyl disilazane sodium salt [1070-89-9]) is obtained, and evolution of NH<sub>2</sub> ceased (2-4 hours). Note that repeated purging with argon or N<sub>2</sub> reduces the reaction time. The condenser is removed and under argon or N<sub>2</sub>, (methoxycarbonylmethylene)triphenylphosphorane (167g, 0.5mol, [2605-67-6]) is added in large portions, and the mixture is again heated (70-80° bath temperature) with stirring for 24-30 hours until the colour of the solution assumes a bright yellow hue [and the IR band of a film of aliquots between KBr plates, with  $\nu_{\max}$  at 1616 cm<sup>-1</sup> of the starting ylide has disappeared, and is replaced by a strong band of the product with  $\nu_{\max}$  at 2090 cm<sup>-1</sup> of constant intensity, indicating completion of reaction]. While still hot, and under argon or N<sub>2</sub>, the mixture is filtered rapidly [through a 20/40 standard taper vacuum filtration adapter and a 350ml fritted funnel of coarse porosity open to the atmosphere through a 2 cm thick layer of basic Al<sub>2</sub>O<sub>3</sub> (III) and a 1cm Celite pad on top, and occasional scraping the top of the pad] into a 2L 24/40 single necked flask. The filtration requires 30-60 minutes and the funnel temperature should be kept between 30-50° to avoid crystallisation. The filtrate is flushed with inert gas and evaporated (rotavap) to dryness. The residue is recrystallised from hot dry toluene (1g/5ml) followed by cooling at -20° overnight. While the mixture is cold, the solids are filtered off, washed with cold toluene (3 x 100ml at 0°). The combined filtrates are evaporated in the same way at low temperature as before to give a second crop. The combined crops are dried to constant weight at high vacuum (0.01mm) to give pure (*triphenylphosphoranylidene*) ketene (106.8-114.9g, 74-76% and ~99.97% pure by <sup>1</sup>H NMR) as a pale yellow, flaky powder with m 173°. It is remarkably stable, it can be stored at ~25° under argon or N<sub>2</sub> for months and can be handled at ~25° without obvious deterioration. It has IR (KBr) with  $\nu_{\max}$  at 2090 (s), 1625 (m), 1436 (m) and 1110 (m) cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) has  $\delta$  at 7.44-7.54 (m, 6H), 7.55-7.63 (m, 3H) and 7.64-7.77 (m, 6H); the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta$  at -10.5 (d, <sup>3</sup>J<sub>P-C</sub> = 185.4 Hz, C $\alpha$ ), 128.8 (d, <sup>3</sup>J<sub>P-C</sub> = 12.9 Hz, C-ortho), 129.6 (d, <sup>1</sup>J<sub>P-C</sub> = 98.5 Hz, C-*ipso*), 132.2 (s, C-*para*), 132.3 (s, C-*meta*) and 145.6 (d, <sup>2</sup>J<sub>P-C</sub> = 63.0 Hz, C $\beta$ ); and the <sup>31</sup>P NMR (162MHz, CDCl<sub>3</sub>) has  $\delta$  at 6.0. Possible detectable impurities are starting ester ylide and toluene detected by <sup>1</sup>H NMR spectroscopy which have singlets at  $\delta$  3.60 and 2.36 respectively. It is a nucleophilic-only C<sub>2</sub>-building block that reacts with a range of electrophiles in a variety of ways. It also undergoes cycloaddition reactions at the polar P-C $\alpha$  bond as well as at the C $\alpha$ =C $\beta$  double bond, and multi-component and domino reactions. [Schobert *Org Synth* **82** 140 2005.]

**Triphenyl silane** [789-25-3] M 260.4, m 45°, b 148-151°/1mm. Purify it by recrystallisation from MeOH. [Gilman & Zuech *J Am Chem Soc* **81** 5925 1959, Westermarck *Acta Chem Scand* **9** 947 1955, IR: Kaplan *J Am Chem Soc* **76** 5880 1954, *Beilstein* **16** II 605, **16** III 1199, **16** IV 1369.]

**Triphenylsilanol (hydroxytriphenylsilane)** [791-31-1] M 276.4, m 150-153°, 151-153°, 154-155°, 156°. It is purified by dissolving in petroleum ether, passing through an Al<sub>2</sub>O<sub>3</sub> column, eluting thoroughly with CCl<sub>4</sub> to remove impurities and then eluting the silanol with MeOH. Evaporation gives crystals with m 153-155°. It can be recrystallised from petroleum ether, CCl<sub>4</sub> or from \*benzene or Et<sub>2</sub>O/petroleum ether (1:1). It has also been recrystallised by partial freezing from the melt to constant melting point. [George & Gilman *J Am Chem Soc* **81** 3288 1959, IR: Tatlock & Rochow *J Org Chem* **17** 1555 1952 and Richards & Thompson *J Chem Soc* 124 1949, *Beilstein* **16** IV 1480.]

**Triphenyl vinyl silane** [18666-68-7] M 286.5, m 58-59°, 57-59.5°, 67-68°, b 190-210°/3mm. The vinylsilane has been recrystallised from EtOH, 95% EtOH, EtOH/\*C<sub>6</sub>H<sub>6</sub>, petroleum ether (b 30-60°) and Et<sub>2</sub>O, and has been distilled under reduced pressure. [Cason & Brooks *J Am Chem Soc* **74** 4582 1952, Nagel & Post *J Org Chem* **17** 1379 1952, *Beilstein* **16** IV 1371.]

**Tri(*n*-propyl) borate** [688-71-1] M 188.1, b 64°/9mm, 175-177°/atm, d<sub>4</sub><sup>20</sup> 0.857, n<sub>D</sub><sup>20</sup> 1.395. Dry the ester over sodium and then distil it, preferably in a vacuum. (cf tributyl borate.) [Charnley et al. *J Chem Soc* 2288 1952, *Beilstein* **1** IV 1436.]

**Tri(*iso*-propyl) borate [(*iso*-PrO)<sub>3</sub>B, boric acid tri-*iso*-propyl ester, Boron *iso*-propoxide]** [5419-55-6] M 188.1, b 75°/76mm, 90°/120mm, 139-141°/760mm, d<sub>4</sub><sup>25</sup> 0.815, n<sub>D</sub><sup>20</sup> 1.376. The borate ester is prepared in 85% yield from *iso*-propanol (250ml) and NaBH<sub>4</sub> (6.84g) followed by dropwise addition of AcOH (0.18g) over a period of 11 minutes, and then refluxing for 4 hours (fume cupboard and a “Dry-ice” condenser as 16.15L of H<sub>2</sub> are released, CARE due to its flammability). Then fractionate the ester through a Widmer column. [Brown et al. *J Am Chem Soc* **78** 3613 1956.] The ester is a good reagent for borylation, e.g. *ortho*-borylating 1-substituted naphthalenes, by reacting with 1-halo- or cyano- naphthalenes (after treatment with LiTMP) to form 2-boryl esters which are later used in Pd-catalysed cross-coupling reactions [Lysén et al. *Synthesis* 3478 2006]. [*Beilstein* **1** H 363, **1** II 382, **1** III 1468, **1** IV 1488.]

**Tri(quinol-8-yl) phosphate** [52429-99-9] M 479.4, m 193-197°, 202-203°. Purify the phosphate by recrystallisation from dimethylformamide. The purity is checked by paper chromatography, R<sub>F</sub> 0.90 [*i*-PrOH/saturated (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>/H<sub>2</sub>O, 2:79:19 as eluent], its IR (KBr) has ν<sub>max</sub> at 1620, 1570 (C=C, C=N) and 1253 (P=O) cm<sup>-1</sup>. [Takaku et al. *Bull Chem Soc Jpn* **47** 779 1974.]

**Tri-*o*-tolylphosphine** [6163-58-2] M 304.4, m 129-130°, pK<sub>Est</sub> ~1.0. Like PPh<sub>3</sub>, it crystallises from hexane, MeOH, diethyl ether, CH<sub>2</sub>Cl<sub>2</sub>/hexane or 95% EtOH. Dry it at 65°/1mm over CaSO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub>. Purify further by chromatography through alumina using (4:1) \*benzene/CHCl<sub>3</sub> as eluent. Of the phosphine ligands studied, P(*o*-Tol)<sub>3</sub> (6 mol%) was the best in the Ru<sub>3</sub>(CO)<sub>12</sub> (2 mol%) catalysed *N*-alkylation of primary amines and secondary alcohols (e.g. hexylamine with 1-phenylethanol to give *N*-octyl 1-phenethylamine) in high yields at 90-110° [Tillack et al. *Tetrahedron Lett* **47** 8881 2006]. The Heck reaction between aryl halides and *n*-butyl acrylates to provide substituted styrenes in ~80 to >90% yields at ~120° is catalysed by Pd(OAc)<sub>2</sub> (2 mol%) and P(*o*-Tol)<sub>3</sub> (8 mol%) without extensive P-C bond cleavage [Herrmann et al. *J Mol Catal A: Chem* **103** 133 1995]. [*Beilstein* **16** III 835.]

**Tris(2-biphenyl) phosphate** [132-28-5] M 554.6, m 115.5-117.5°. Crystallise it from MeOH containing a little acetone. (cf triphenyl phosphate.)

**Tris(2,4-di-*tert*-butylphenyl)phosphite** [31570-04-4] M 646.9, m 181-184°. If the ester is suspect, e.g. partly hydrolysed, dissolve it in Et<sub>2</sub>O, wash the solution successively with aqueous 5% NaOH, H<sub>2</sub>O and saturated aqueous NaCl, then shake with charcoal, filter, dry with Na<sub>2</sub>SO<sub>4</sub>, filter again, evaporate and add petroleum ether when the volume has decreased considerably to crystallise the ester. [Verkade & Coskren in *Organophosphorus*

*Compounds* (Kosolapoff & Maier eds) Wiley Vol 6 pp 211-577 1973.] It is used as a processing stabiliser for polymers.

**Tris(1,2-dioxyphenyl)cyclotriphosphazine** {trispiro[1,3,5,2,4,6-triazatriphosphorine]-2,2':-2,4'':2,6'''-tris(1,3,2)benzodioxaphosphole} [311-03-5] **M 459.0, m 244-245°, 245°, 245-246°**. Recrystallise this phosphazine from \*C<sub>6</sub>H<sub>6</sub> or chlorobenzene, then triple sublime it (175°/0.1mm, 200°/0.1mm, 230°/0.05mm). Its UV has  $\lambda_{\max}$  nm (log  $\epsilon$ ) at 276 (3.72), 271 (3.79) 266sh (3.68) and 209 (4.38) in MeCN. Its IR has  $\nu_{\max}$  at 1270 (O-Ph), 1220 (P=N), 835 (P-O-Ph) and 745 (Ph) cm<sup>-1</sup>. [Alcock *J Am Chem Soc* **86** 2591 1964, Alcock et al. *J Am Chem Soc* **98** 5120 1976, Meirovitch *J Phys Chem* **88** 1522 1984.]

**(±)-Tris(2-ethylhexyl) phosphate (TEHP, tri-isooctylphosphate, "trioctyl" phosphate, [78-42-2, 25103-23-5]** **M 434.6, b 186°/1mm, 219°/5mm, d<sup>25</sup> 0.92042, n<sub>D</sub><sup>20</sup> 1.44464**. TEHP, in an equal volume of diethyl ether, is shaken with aqueous 5% HCl, and the organic phase is filtered to remove traces of pyridine (used as a solvent during manufacture) as its hydrochloride. This layer is shaken with aqueous Na<sub>2</sub>CO<sub>3</sub>, then water, and the ether is distilled off at room temperature. The ester is then filtered, dried for 12 hours at 100°/15mm, and again filtered, then shaken intermittently for 2 days with activated alumina (100g/L). It is decanted through a fine sintered-glass disc (with exclusion of moisture), and distilled under vacuum. [French & Muggleton *J Chem Soc* 5064 1957.] \*Benzene can be used as a solvent (to give 0.4M solution) instead of ether. Its IR has  $\nu_{\max}$  at 1702, 1701, 481 and 478cm<sup>-1</sup> [Bellamy & Becker *J Chem Soc* 475 1952]. The *uranyl nitrate* salt is purified by partial crystallisation from hexane [Siddall & Dukes *J Am Chem Soc* **81** 790 1959, Siddall *J Am Chem Soc* **81** 4176 1959]. [Beilstein 1 IV 1786.]

**Tris(hydroxymethyl)phosphine (phosphinidynetrimethanol) [2767-80-8]** **M 121.1, m 48-56°, 55°, b 111-113°/2.5mm, n<sub>D</sub><sup>20</sup> 1.5497, pK 5.5 (half neutralisation point, glass electrode)**. It is obtained in two ways by neutralising precisely the THPC (Tetrakis(hydroxymethyl)phosphonium chloride, [124-64-1]). In the *first* method, NaOH (1.6g, 40mmol) in H<sub>2</sub>O (25ml) is added rapidly to THPC (7.6g, 40mmol, see above [124-64-1]) in H<sub>2</sub>O (50ml) under N<sub>2</sub> at room temperature in a closed system connected to a glass buret. No evolution of gas should be observed on stirring for 20 hours indicating that the stoichiometric amount of alkali has been added. Any liberation of gas (H<sub>2</sub>) would indicate that excess of alkali had been added which would liberate H<sub>2</sub> with the formation of the corresponding phosphine oxide (see next entry, and Hoffman *J Am Chem Soc* **52** 2995 1930). The solvent is then removed *in vacuo* at 50-60°, EtOH (under N<sub>2</sub>) is added, the solid NaCl is filtered off, and the EtOH is removed *in vacuo* (45°) to give the *free base* (4.8g, 97%) as a clear viscous liquid which can be distilled at high vacuum in an atmosphere of N<sub>2</sub> (N<sub>2</sub> bleed). An IR (film) should indicate the absence of P=O and the presence of primary OH and P-CH<sub>2</sub> bands [Grayson *J Am Chem Soc* **85** 79 1963].

In the *second* method, under N<sub>2</sub> flushing and stirring, a solution of THPC (25g, 131mmol) in distilled H<sub>2</sub>O (100ml), into which is immersed a glass pH electrode, is treated with ~ 150ml (0.21 equivalents) of 20-50 mesh Dowex-1 x 8 resin (in the OH<sup>-</sup> form) during 45 minutes until the pH was ~8.3-8.5. The resin is filtered off, washed with H<sub>2</sub>O, the combined aqueous solutions are evaporated (rotavap) at 65-70°, then under high vacuum for 30 minutes, then further at -30° to give the *free base* as a viscous liquid which formed a waxy solid. Its IR (film between NaCl plates) should have no bands at  $\nu_{\max}$  1052 (of THPC), or 1043 or 1135 (for P=O), but intense broad absorption at 1010 cm<sup>-1</sup>. The <sup>1</sup>H NMR (60MHz, D<sub>2</sub>O, internal TMS, or DSS at 0) has  $\delta$  for CH<sub>2</sub> at 4.15 ( $J_{\text{PCH}} = 5.5\text{Hz}$ ,  $J_{\text{13CH}} = \sim 148\text{Hz}$ ); and the <sup>31</sup>P NMR (40MHz, D<sub>2</sub>O, external 85% H<sub>3</sub>PO<sub>4</sub>) has  $\delta$  at -25.8. [Ellzey et al. *J Org Chem* **37** 3453 1972.] With *n*-butyl iodide it formed *n*-butyl phosphonium iodide which gave, on treatment with sodium tetraphenylboron, the *n*-butyl tris(hydroxymethyl)phosphonium tetraphenylboron complex, **m 145-146°(dec)** after drying and recrystallising from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>; and the *methiodide* gave with Ph<sub>4</sub>BNa, the *methyl tris(hydroxymethyl)phosphonium tetraphenylboron complex*, **m 170-171°(dec)** after drying and recrystallising from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub> [Grayson *J Am Chem Soc* **85** 79 1963.]

**Tris(hydroxymethyl)phosphine oxide [1067-12-5]** **M 137.1, m 50-52°, 54-55° (69° and 70° have also been reported)**. The oxide is prepared by slowly adding about 130% excess (over an equimolar amount) of 20-50 mesh Dowex-1 x 8 resin (in the OH<sup>-</sup> form) to an aqueous solution of tris(hydroxymethyl)phosphine (25.0g, see preceding entry) which led to vigorous evolution of H<sub>2</sub>. After standing overnight the resin is filtered off, and the filtrate is evaporated (rotavap, vacuum) to give the *oxide* (16.1g, 88%) which is recrystallised from absolute EtOH to yield hygroscopic crystals (11.9g). Store in a dry atmosphere. The IR has strong bands at  $\nu_{\max}$  1043

and  $1134\text{ cm}^{-1}$ ; the  $^1\text{H}$  NMR (60MHz,  $\text{D}_2\text{O}$ , internal TMS, or DSS at 0) has  $\delta$  for  $\text{CH}_2$  at 4.20 ( $J_{\text{PCH}} = 3.1\text{Hz}$ ,  $J_{^{13}\text{CH}} = \sim 146\text{Hz}$ ); and the  $^{31}\text{P}$  NMR (40MHz,  $\text{D}_2\text{O}$ , external 85%  $\text{H}_3\text{PO}_4$ ) has  $\delta$  at -48.7. [Ellzey et al. *J Org Chem* **37** 3453 1972, Anteunis et al. *Bull Soc Chim Belg* **74** 622 1965.] The *tribenzoate*, prepared by boiling THPC with excess of aqueous NaOH solution until no further evolution of  $\text{H}_2$  occurred followed by treatment with a slight excess of  $\text{PhCOCl}$ , gave fine needles **m 111**<sup>o</sup> upon recrystallisation from MeOH [Hoffman *J Am Chem Soc* **43** 1684 1921].

**Trisodium 8-hydroxy-1,3,6-pyrenetrisulfonate (Pyranine Solvent Green 7)** [6358-69-6] **M 524.4, m >300(dec), CI 59040,  $\lambda_{\text{max}}$  403nm**. Purify the salt by chromatography through an alumina column, and elute with *n*-propanol/water (3:1, v/v). Recrystallise it from aqueous acetone (5:95, v/v) using decolorising charcoal. [Beilstein **1** III 565.] **IRRITANT**.

**Trisodium 1,3,6-naphthalenetrisulfonate** [5182-30-9] **M 434.2**. The *free acid* is obtained by passing the salt through an ion-exchange column and converting it to the lanthanum salt by treatment with  $\text{La}_2\text{O}_3$ . This salt is crystallised twice from hot water. [The much lower solubility of  $\text{La}_2(\text{SO}_4)_3$  and its retrograde temperature dependence allows a good separation from sulfate impurity]. The lanthanum salt is then passed through an appropriate ion-exchange column to obtain the free acid, the sodium or potassium salt. (The sodium salt is *hygroscopic*.) [Atkinson et al. *J Am Chem Soc* **83** 1570 1961.] It can also be recrystallised from aqueous acetone [Okahata et al. *J Am Chem Soc* **108** 2863 1986].

**Tris(2,2,2-trifluoroethyl) phosphite** [370-69-4] **M 328.1, b 130-131<sup>o</sup>/743mm,  $d_4^{20}$  1.487,  $n_D^{20}$  1.324**. Fractionate the phosphite through a 10 inch Helipak column [Krogh et al. *J Org Chem* **19** 1124 1954].

**Tris(trimethylsilyl)silane (TTMSS)** [1873-77-4] **M 248.7, b 73<sup>o</sup>/5mm,  $d_4^{20}$  0.808,  $n_D^{20}$  1.49**. Purify it by fractional distillation and taking the middle cut. Store it under  $\text{N}_2$  or Ar as it is **PYROPHORIC** and is an **IRRITANT**. [Chatgililoglu et al. *J Org Chem* **53** 3641 1988, Balestri et al. *J Org Chem* **56** 678 1991, Chatgililoglu *Acc Chem Res* **25** 188 1992, NMR: Gilman et al. *J Organomet Chem* **4** 163 1965.]

**Tri(*p*-tolyl) phosphate** [20756-92-7, 1330-78-5 (*isomeric tritolyl phosphate mixture*)] **M 368.4, m 77<sup>o</sup>, b 232-234<sup>o</sup>,  $d_4^{25}$  1.16484,  $n_D^{20}$  1.56703**. Dry the ester with  $\text{CaCl}_2$ , percolate it through a column of alumina, then distil it under a vacuum. *Alternatively*, pass it through a packed column of alumina at 150<sup>o</sup>, with a counter-current stream of nitrogen, under reduced pressure, to remove residual traces of volatile impurities. It also crystallises from petroleum ether (b 60-80<sup>o</sup>). [Cherbuliez in *Organic Phosphorus Compounds* (Kosolapoff & Maier eds) Wiley-Interscience Vol **6** pp 454-457 1973.]

**Tropaeolin 00 {4-[(4-anilinophenyl)azo]benzenesulfonic acid Na salt}** [554-73-4] **M 375.4,  $\text{pK}_{\text{Est}(1)} \sim 2.3$ ,  $\text{pK}_{\text{Est}(2)} \sim 5.8$ ,  $\text{pK}_{\text{Est}(3)} \sim 10.3$** . Recrystallise it twice from water [Kolthoff & Gus *J Am Chem Soc* **60** 2516 1938]. It is an indicator which is red at pH 1.4 and orange-yellow at pH 3.2. [Beilstein **16** II 171.]

**Vinyl chlorosilane** [75-94-5] **M 161.5, b 17.7<sup>o</sup>/46.3mm, 82.9<sup>o</sup>/599.4mm, 92<sup>o</sup>/742mm, 91-91.5<sup>o</sup>/atm,  $d_4^{20}$  0.1.2717,  $n_D^{20}$  1.435**. Fractionally distil the chlorosilane at atmospheric pressure. It is water sensitive and is stored in the dark and it is likely to polymerise. [Müller & Schnurrbusch *Chem Ber* **91** 1805 1958, Munkelt & Müller *Chem Ber* **92** 1012 1959, Polarography: Abrahamson & Reynolds *Anal Chem* **24** 1827 1952, Beilstein **4** IV 4258.]

**3-Vinylphenylboronic acid ([3-ethenylphenyl]-dihydroxyborane)** [15016-43-0] **M148.0, m 141-147<sup>o</sup>, 145-160<sup>o</sup>,  $\text{pK}_{\text{Est}(1)} \sim 8.8$** . This binder for paper or glass crystallises in white plates from  $\text{H}_2\text{O}$  (**m** 144-145<sup>o</sup>), and its IR has  $\nu_{\text{max}}$  at 919 and 995 ( $\text{CH}_2\text{CH}=\text{CH}_2$ ), 1350 (B-O) and 3220 (OH)  $\text{cm}^{-1}$ . The *dibromo* derivative crystallises in white needles from  $\text{H}_2\text{O}$  with **m** 196-197.5<sup>o</sup>. [Wesley & Rush *J Org Chem* **27** 2598 1962, Beilstein **16** III 1279, **16** IV 1677.]

**Vinylphosphonic acid** [1746-03-8] **M 108.0**, **m 41-45°**, **d**<sup>25</sup> **1.389**, **pK**<sub>1</sub><sup>20</sup> **3.48**, **pK**<sub>2</sub><sup>20</sup> **8.54** (50% EtOH). This fireproofing agent, and ingredient for making polymers, is obtained as a syrup on hydrolysing vinylphosphonyl dichloride with cold H<sub>2</sub>O and solidifies on prolonged drying over P<sub>2</sub>O<sub>5</sub>/KOH. When distilled at 235-240°/0.0006mm, it gives the *anhydride* (**d**<sub>4</sub><sup>20</sup> 1.304, **n**<sub>D</sub><sup>20</sup> 1.5874) [Kabachnik & Medvedi *Izvest Akad Nauk SSSR, Ser Khim* 868 1953, *Chem Abstr* 54 10834 1960]. It is best kept as the *sodium salt* (**m** 350°) which precipitates when a solution of EtOH containing NaOEt (from 2g of Na) is added to vinylphosphonic acid (3.2g), and is recrystallised from EtOH (5.1g, quantitative). The *p-anisidinium salt* forms mauve prisms **m** 250° (from EtOH/Et<sub>2</sub>O). The *dimethyl ester*, [4645-32-3] **M 136.1**, **d**<sub>4</sub><sup>20</sup> 1.1405, **n**<sub>D</sub><sup>20</sup> 1.4330, has **b** 72.5°/10mm and 197-202°/760mm. [Kabachnik et al. *J Gen Chem USSR (Engl Trans)* 33 375 1963, *Beilstein* 4 IV 3568.]

**Xylenol Orange (sodium salt)**. See entry in "Heterocyclic Compounds" in this Chapter.

**Zinc diethyldithiocarbamate** [14324-55-1] **M 561.7**, **m 178-181°**, **pK**<sup>25</sup> **3.04** (for Et<sub>2</sub>NCS<sub>2</sub><sup>-</sup>). Crystallise this herbicide several times from hot toluene or from hot CHCl<sub>3</sub> by addition of EtOH. It also crystallises from xylene, **m** 180°. [*Beilstein* 4 II 613.] **TOXIC**.

**Zinc dimethyldithiocarbamate** [137-30-4] **M 305.8**, **m 151-152°, 248-250°**, **pK**<sup>25</sup> **3.36** (for Me<sub>2</sub>NCS<sub>2</sub><sup>-</sup>). Crystallise this herbicide several times from hot toluene or from hot CHCl<sub>3</sub> by addition of EtOH. [*Beilstein* 4 III 149, 4 IV 234.]

**Zinc ethylenebis(dithiocarbamate) (Zineb)** [12122-67-7] **M 249.7**. Crystallise this herbicide several times from hot toluene or from hot CHCl<sub>3</sub> by addition of EtOH. It is a skin irritant. [*Beilstein* 4 III 149, 4 IV 234.]

**Zinc phenol-*o*-sulfonate (8H<sub>2</sub>O) (Phenozin)** [127-82-2] **M 555.8**. Phenozin crystallises from warm water by cooling to 0°. It effloresces in dry air, and loses all its H<sub>2</sub>O at 120°. Its solubility in H<sub>2</sub>O is 63% at 20° and 250% at boiling point, and in EtOH it is 55% at 50°. [*Beilstein* 11 H 53, 11 I 234, 11 IV 574.]

**Zinc trifluoromethanesulfonate** [54010-75-2] **M 363.5**, **m >300°**. This zinc salt should be dried at 125° for 2 hours at 3mm before use. It is soluble in CH<sub>2</sub>Cl<sub>2</sub> but insoluble in petroleum ether. [Corey & Shimoji *Tetrahedron Lett* 24 169 1983.]

# CHAPTER 5

## PURIFICATION OF INORGANIC AND METAL-ORGANIC CHEMICALS

### INTRODUCTION

The most common method of purification of inorganic species is by recrystallisation, usually from water. However, especially with salts of weak acids or of cations other than the alkaline and alkaline earth metals, care must be taken to minimise hydrolysis. This can be achieved, for example, by recrystallising acetates in the presence of dilute acetic acid. Nevertheless, there are many inorganic chemicals that are too insoluble or are hydrolysed by water so that no general purification method can be given. It is convenient that many inorganic substances have large temperature coefficients for their solubility in water, but in other cases recrystallisation is still possible by partial solvent evaporation.

Organo-metallic compounds, on the other hand, behave very much like organic compounds, e.g. they can be redistilled and may be soluble in organic solvents. A note of **caution** should be made about handling organo-metallic compounds, e.g. arsines, because of their **potential toxicities**, particularly when they are volatile. Generally the suppliers of such compounds provide details about their safe manipulation. These should be read carefully and adhered to closely. If in any doubt, always assume that the materials are lethal and treat them with utmost care. The same **safety precautions** about the handling of substances as stated in Chapter 4 should be followed here (see Chapter 1).

For information on **ionization (pK)** see Chapter 1, pp 34-36, and Chapter 4, p 103. In order to avoid repetition, the literature (or predicted) pK values of anionic and/or cationic species are usually reported at least once, and in several cases are entered for the free acid or free base; e.g.  $\text{Na}_2\text{SO}_4$  will have a pK value for  $\text{Na}^+$  at the entry for NaOH and the pK values for  $\text{SO}_4^{2-}$  at the entries for  $\text{H}_2\text{SO}_4$ . When the pK values of the organic counterions are not given in this chapter, as in the case of sodium benzoate, the reader is referred to the value(s) in Chapter 4, Aromatic Compounds, e.g. of benzoic acid.

**Abbreviations** of titles of periodicals are defined as in the Chemical Abstracts Service Source Index (CASSI).

**Benzene**, which has been used as a solvent successfully and extensively in the past for reactions and purification by chromatography and crystallization, is now considered a **very dangerous substance**, so it has to be used with extreme care. We emphasised that an alternative solvent to benzene (e.g. toluene, toluene-petroleum ether, or a petroleum ether to name a few) should be used first. However, if benzene has to be used, then all operations have to be performed in well-ventilated fumehoods and precautions taken to avoid inhalation and contact with skin and eyes. Whenever benzene is mentioned in the text, an asterisk, e.g.  $\text{*C}_6\text{H}_6$  or  $\text{*benzene}$ , is inserted to remind the user that special precaution should be adopted.

**Organic dyes** which are *not* complexed, or are salts of metals, are included in Chapter 4 (use the CAS Registry Numbers to find them). Commercially available polymer-supported reagents are indicated with § under the appropriate reagent.

This chapter is subdivided into two sections: the **Purification of Inorganic Compounds** and the **Purification of Metal-Organic Compounds** which includes ammonium and metal salts of organic acids.

## INORGANIC COMPOUNDS

**Alumina (aluminium oxide) (neutral)** [1344-28-1] **M 102.0 (anhydrous)**. Stir the oxide with hot 2M HNO<sub>3</sub>, either on a steam bath for 12 hours (changing the acid every hour) or three times for 30 minutes, then wash it with hot distilled water until the washings have pH 4, and follow by three washings with hot MeOH. The product is dried at 270° [Angyal & Young *J Am Chem Soc* **81** 5251 1959]. For the preparation of alumina for chromatography see Chapter 1. [For α, β and γ Al<sub>2</sub>O<sub>3</sub> see Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 822-823 1963 and Wagner in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1662 1965.]

**Aluminium ammonium sulfate decahydrate** [7784-26-1] **M 453.3, m 93°, m 95°, pK<sub>1</sub><sup>25</sup> 4.89, pK<sub>2</sub><sup>25</sup> 5.43, pK<sub>3</sub><sup>25</sup> 5.86 (Al<sup>3+</sup> aquo), pK<sub>4</sub><sup>25</sup> 11.22 [aluminate Al(OH)<sub>4</sub><sup>-</sup>]**. Crystallise it from hot H<sub>2</sub>O and cool in ice. When the melt is heated, it loses NH<sub>3</sub> and H<sub>2</sub>SO<sub>4</sub>, and gives pure alumina at red heat. Solubility (%) in H<sub>2</sub>O is 3.9% (0°), 15.0 (20°) and 135 (100°).

**Aluminium bromide** [7727-15-3] **M 266.7, m 97°, b 114°/10mm, d<sub>4</sub><sup>18</sup> 3.205**. Reflux it and then distil it from pure aluminium chips in a stream of nitrogen into a flask containing more of the chips. It is then redistilled under vacuum into ampoules [Tipper & Walker *J Chem Soc* 1352 1959]. Anhydrous conditions are essential, and the white to very light brown solid distillate can be broken into lumps in a dry-box (under nitrogen). It fumes in moist air. [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 812-813 1963.]

**Aluminium caesium sulfate dodecahydrate** [7784-17-0 (12H<sub>2</sub>O), 14284-36-7] **M 568.2**. Recrystallise it from hot water (3ml/g).

**Aluminium chloride (anhydrous)** [7446-70-0] **M 133.3, m 192.6°, d<sub>4</sub><sup>17</sup> 2.465**. Sublime it several times in an all-glass system under nitrogen at 30-50mm pressure. It has also been sublimed in a stream of dry HCl and has been subjected to a preliminary sublimation through a section of granular aluminium metal [for manipulative details see Jensen *J Am Chem Soc* **79** 1226 1957]. It fumes in moist air. Used in Friedel-Crafts reaction.

**Aluminium fluoride (anhydrous)** [7784-18-4] **M 84.0, m 250°**. The technical material may contain up to 15% alumina, and minor impurities such as aluminium sulfate, cryolite, silica and iron oxide. Reagent grade AlF<sub>3</sub> (hydrated) contains only traces of impurities, but its water content is variable (and may be up to 40%). It can be dried by calcining at 600-800° in a stream of dry air (some hydrolysis occurs), followed by vacuum distillation at low pressure in a graphite system, heated to approximately 925° (condenser at 900°) [Henry & Dreisbach *J Am Chem Soc* **81** 5274 1959]. Its solubility in H<sub>2</sub>O is 0.5%. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 225-226 1963.]

**Aluminium nitrate nonahydrate** [7784-27-2 (9H<sub>2</sub>O); 13473-90-0] **M 375.1**. Crystallise the nitrate from dilute HNO<sub>3</sub>, and dry it by passing dry nitrogen through the crystals for several hours at 40°. After 2 recrystallisations of ACS grade, it had S, Na and Fe at 2.2, 0.01 and 0.02 ppm, respectively.

**Aluminium potassium sulfate dodecahydrate (alum)** [7784-24-9] **M 474.4, m 92°**. Crystallise it from weak aqueous H<sub>2</sub>SO<sub>4</sub> (ca 0.5ml/g). Its solubility (%) in H<sub>2</sub>O is 5.7 (0°), 12.0 (20°) and 136.9 (100°).

**Aluminium rubidium sulfate dodecahydrate** [7784-29-4] **M 496.2**. Crystallise the double salt from aqueous H<sub>2</sub>SO<sub>4</sub> (ca 2.5ml/g).

**Aluminium sulfate (anhydrous)** [10043-01-3] **M 342.2, m 765°(dec), Al<sub>2</sub>O<sub>3</sub> 14-18 H<sub>2</sub>O [17927-65-0], Al<sub>2</sub>O<sub>3</sub> 18 H<sub>2</sub>O [7784-31-8]**. It crystallises from hot dilute H<sub>2</sub>SO<sub>4</sub> (1 ml/g) on cooling in ice. When a solution of alumina (Al<sub>2</sub>O<sub>3</sub>) in conc H<sub>2</sub>SO<sub>4</sub> is slowly cooled, Al<sub>2</sub>SO<sub>4</sub> 17 or 18H<sub>2</sub>O deposits as a crystalline mass. Al<sub>2</sub>SO<sub>4</sub> 17H<sub>2</sub>O is the stable form in equilibrium with its saturated aqueous solution at 25° [Smith *J Am Chem Soc* **64** 41

1942]. This is purified by dissolving it in a small volume of H<sub>2</sub>O and adding EtOH until the sulfate readily crystallises from the oily supersaturated solution. It forms Al<sub>2</sub>O<sub>3</sub>·16H<sub>2</sub>O between 0–112°. On gradual heating, the hydrate melts, giving the anhydrous salt at ca 250°. Several hydrates up to 27H<sub>2</sub>O have been described. Further heating to red heat (~ 600–800°) causes decomposition to Al<sub>2</sub>O<sub>3</sub> + SO<sub>3</sub> + SO<sub>2</sub> and O<sub>2</sub> [Cobb *J Soc Chem Ind* **29** 250 1910]. The ACS reagent is Al<sub>2</sub>O<sub>3</sub>·18H<sub>2</sub>O (98+%).

**Ammonia (gas)** [7664-41-7] **M 17.0, pK<sup>25</sup> 9.25**. Major contaminants are water, oil and non-condensable gases. Most of these impurities are removed by passing the ammonia through a trap at -22° and condensing it at -176° under vacuum. Water is removed by distilling the ammonia into a tube containing a small lump of sodium. Also dry it by passage through porous BaO, or over alumina followed by glass wool impregnated with sodium (prepared by soaking the glass wool in a solution of sodium in liquid ammonia and evaporating off the ammonia). It can be rendered oxygen-free by passage through a solution of potassium in liquid ammonia. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 460-463 1963.] **AMMONIA (gas, liquid or aqueous solution) is very irritating and should not be inhaled in any quantity as it can lead to olfactory paralysis (temporary or permanent).**

**Ammonia (liquid)** [7664-41-7] **M 17.0, m -77.8°, b -33.35°, n<sub>D</sub><sup>20</sup> 1.325, d<sub>4</sub><sup>20</sup> 0.597, d<sup>-79</sup> 0.817g/ml**. Dry the liquid, and store it, with sodium in a steel cylinder, then distil and condense it by means of liquid air, the non-condensable gases being pumped off. In order to obtain liquid NH<sub>3</sub> from a cylinder, turn the cylinder upside-down (i.e. with the valve at the bottom, use a metal stand to secure it in this position; a special stand can be constructed for it) and lead a plastic tube from the tap to a measuring cylinder placed in an efficient fume cupboard which is kept running. Turn the tap on and allow the ammonia to be released. At first, gas and liquid will splatter out (make sure that the plastic tube is secure), but soon the liquid will drip into the measuring cylinder. The high latent heat of evaporation will cool the ammonia so that the liquid will remain cool and not boil vigorously. If the ammonia is required dry, the necessary precautions should be taken, i.e. the gas is allowed to flow through tubes packed with coarse CaO pellets. **AMMONIA (gas, liquid or aqueous solution) is very irritating and should not be inhaled in any quantity as it can lead to olfactory paralysis (temporary or permanent).**

**Ammonia (aqueous)** [7664-41-7] **M 17.0 + H<sub>2</sub>O, d<sub>4</sub><sup>20</sup> 0.90 (saturated, 27% w/v, 14.3 N), pK<sup>25</sup> 9.25 (pK<sub>b</sub> at 25° = 4.75, i.e. 14.00-9.25, and K<sub>b</sub> = 1.81 x 10<sup>-5</sup>)**. Obtained metal-free by saturating distilled water, in a cooling bath, with ammonia (from a cylinder) gas. *Alternatively*, isothermal distillation can be used by placing a dish of concentrated aqueous ammonia and a dish of pure water in an empty desiccator and leaving to equilibrate for several days. **AMMONIA (gas, liquid or aqueous solution) is very irritating and should not be inhaled in any quantity as it can lead to olfactory paralysis (temporary or permanent).**

**Ammonium bisulfate (ammonium hydrogen sulfate (NH<sub>4</sub>)<sub>2</sub>HSO<sub>4</sub>)** [7803-63-6] **M 115.1, m ~147°, d<sub>4</sub><sup>20</sup> 1.79, pK<sup>25</sup> 1.96 (HSO<sub>4</sub><sup>-</sup>)**. It crystallises from water at room temperature (1ml/g) on adding EtOH and cooling. (NH<sub>4</sub>)HSO<sub>4</sub> is formed as deliquescent rhombic crystals on cooling a solution of (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> in hot conc H<sub>2</sub>SO<sub>4</sub> but EtOH decomposes it to (NH<sub>4</sub>)<sub>3</sub>H(SO<sub>4</sub>)<sub>2</sub> [Dunncliff *J Chem Soc* **123** 476 1923]. When powdered (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is heated below 100°, it loses NH<sub>3</sub>, and at 300° it is completely converted to fused (NH<sub>4</sub>)HSO<sub>4</sub>, **m** 140°, but >300° it decomposes to SO<sub>2</sub> and N<sub>2</sub> [Smith *J Chem Soc* **30** 253 1911].

**Ammonium bromide** [12124-97-9] **M 98.0, m 450°(sublimes), d<sub>4</sub><sup>20</sup> 2.43**. It crystallises from 95% EtOH and is slightly hygroscopic.

**Ammonium chloride** [12125-02-9] **M 53.5, m 338°(sublime point, without melting), d<sub>4</sub><sup>20</sup> 1.53**. Crystallise it several times from conductivity water (1.5ml/g) between 90° and 0°. It sublimes. The salt is fully ionised in aqueous solution, i.e. K ≈ ∞. A 1M aqueous solution of NH<sub>4</sub>Cl has a pH of ~4.7, i.e. is acidic. After one crystallisation, ACS grade has: metal(ppm) As (1.2), K (1), Sb (7.2), V (10.2). [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 812 1963.]

**Ammonium chromate** [7788-98-9] **M 152.1, m 185°(dec),  $d_4^{20}$  1.81,  $pK_1^{25}$  0.74,  $pK_2^{25}$  6.49 (for  $H_2CrO_4$ ).** Crystallise it from weak aqueous ammonia (ca 2.5ml/g) by cooling from room temperature. It loses  $NH_3$  on heating to form ammonium dichromate. **POISONOUS.**

**Ammonium dichromate** [7789-09-5] **M 252.1, m 170°(dec),  $d_4^{20}$  1.26.** It crystallises from weak aqueous HCl (ca 1ml/g). It decomposes rapidly on heating. (Possible carcinogen and is **POISONOUS**).

**Ammonium dihydrogen arsenate** [13462-93-6] **M 159.0, m 300°(dec).** Crystallise it from water (!ml/g). **POISONOUS.**

**Ammonium dihydrogen orthophosphate** [7722-76-1] **M 115.0, m 190°,  $d_4^{20}$  1.80.** Crystallise it from water (0.7ml/g) between 100° and 0°.

**Ammonium ferric sulfate dodecahydrate** [7783-83-7 (12 $H_2O$ ), 10138-04-2 (anhydrous)] **M 482.2, m ~37°,  $d_4^{20}$  1.71.** Crystallise it from aqueous ethanol.

**Ammonium ferrous sulfate hexahydrate [Mohr's salt]** [7783-85-9 (6 $H_2O$ ), 10045-89-3 (anhydrous)] **M 392.1, m 100°(dec),  $d_4^{20}$  1.86.** A solution in warm water (0.67g/ml) is cooled rapidly to 0°, and the resulting light bluish-green monoclinic crystals are filtered at the pump, washed with cold distilled water and pressed between sheets of filter paper to dry it. The solubility at 25° is 0.36g/ml. It separates as an almost white powder when a saturated aqueous solution is diluted with EtOH.

**Ammonium hexachloroiridate (IV)** [16940-92-4] **M 441.0.** It is precipitated several times from aqueous solution by saturation with ammonium chloride. This removes any palladium and rhodium. It is then washed with ice-cold water and dried over conc  $H_2SO_4$  in a vacuum desiccator. If osmium or ruthenium is present, it can be removed as the tetroxide by heating with conc  $HNO_3$ , followed by conc  $HClO_4$ , until most of the acid has been driven off. (This treatment is repeated.) The near-dry residue is dissolved in a small amount of water and added to excess  $NaHCO_3$  solution and bromine water. On boiling, iridic (but not platinic) hydroxide is precipitated. It is dissolved in HCl and precipitated several times, then dissolved in HBr and treated with  $HNO_3$  and HCl to convert the bromides to chlorides. Saturation with ammonium chloride and cooling precipitates ammonium hexachloroiridate which is filtered off and purified as above [Woo & Yost *J Am Chem Soc* **53** 884 1931].

**Ammonium hexacyanoferrate II hydrate** [14481-29-9] **M 284.1, m dec on heating.** The pale yellow trihydrate powder can be washed with 10% aqueous  $NH_3$ , filtered, then washed several times with EtOH and  $Et_2O$ , and dried at room temperature. It decomposes in a vacuum above 100° and should be stored away from light and under  $N_2$ . In light and air it decomposes by losing  $NH_3$ . [Lux in *Handbook of Preparative Inorganic Chem (Ed. Brauer)* Academic Press Vol II p 1509 1965.]

**Ammonium hexafluorophosphate** [16941-11-0] **M 163.0,  $d_4^{18}$  2.181,  $pK_1^{25}$  ~0.5,  $pK_2^{25}$  5.12 (for fluorophosphoric acid  $H_2PO_3F$ ).** It crystallises from  $H_2O$  in square plates and decomposes on heating before melting. Its solubility in  $H_2O$  at 20° is 74.8% w/v, and it is very soluble in  $Me_2CO$ , MeOH, EtOH and MeOAc, but is decomposed by boiling mineral acids. [Lange & Müller *Chem Ber* **63** 1063 1930, Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 195 1963.]

**Ammonium hexafluorosilicate** [16919-19-0] **M 178.1,  $pK_2$  1.92 (for  $H_2SiF_6$ ).** Crystallise the salt from water (2ml/g). After 3 recrystallisations, the Technical grade salt has Li, Na, K and Fe at 0.3, 0.2, 0.1 and 1.0 ppm respectively.

**Ammonium hypophosphite** [7803-65-8] **M 83.0.** Crystallise it from hot EtOH.

**Ammonium iodate** [13446-09-8] **M 192.9,  $pK^{25}$  0.79 ( $IO_3^{3+}$ ).** Ammonium iodate crystallises from water (8ml/g) on cooling from 100° to 0°.

**Ammonium iodide** [12027-06-4] **M 144.9**, sublimes with dec  $\sim 405^\circ$ ,  $d_4^{20}$  2.51. The iodide crystallises from EtOH on addition of ethyl iodide, and is very *hygroscopic*. Store it in the dark. Its solubility is 177g in 100g of H<sub>2</sub>O at 25°. [Schmeisser in *Handbook of Preparative Inorganic Chem* (Ed. Brauer) Academic Press Vol I p 289-290 1963.]

**Ammonium magnesium chloride hexahydrate** [60314-43-4] **M 256.8**. It crystallises from water (6ml/g) by partial evaporation in a desiccator over KOH (deliquescent).

**Ammonium magnesium sulfate hexahydrate** [20861-69-2] **M 360.6**. It crystallises from water (1ml/g) between 100° and 0°.

**Ammonium manganous sulfate hexahydrate** [13566-22-8] **M 391.3**. It crystallises from water (2ml/g) by partial evaporation in a desiccator.

**Ammonium molybdate** [13106-76-8] **M 196.0**,  $pK_1^{25}$  0.9 (proton addition),  $pK_2^{25}$  3.57,  $pK_3^{25}$  4.08 (for H<sub>2</sub>MoO<sub>4</sub>). Crystallise the salt from water (2.5ml/g) by partial evaporation in a desiccator. When a solution of MoO<sub>3</sub> in excess of hot concentrated NH<sub>3</sub> is cooled, *normal* ammonium molybdate (NH<sub>4</sub>)<sub>2</sub>MoO<sub>4</sub> crystallises out. However, when this solution is made to evaporate which allows the pH to drop to 6, the *common* hydrated ammonium *paramolybdate* (NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub> · 4H<sub>2</sub>O {or 3[(NH<sub>4</sub>)<sub>2</sub>O], 7MoO<sub>3</sub>, 4H<sub>2</sub>O, [12054-85-2] **M 1235.9**} crystallises out. This was an old formula which was confirmed by Sturdivant. [Sturdivant *J Am Chem Soc* **59** 630 1937. Grüttner & Jauder in *Handbook of Preparative Inorganic Chem* (Ed. Brauer) Academic Press Vol II p 1711 1965.]

**Ammonium nickel sulfate hexahydrate** [7785-20-8 (6H<sub>2</sub>O), 15699-18-0 (anhydrous)] **M 395.0**,  $d_4^{20}$  1.923. Crystallise this salt from water (3ml/g) on cooling from 90° to 0°.

**Ammonium nitrate** [6484-52-2] **M 80.0**, **m 165°(moist salt), 210°(dec explosively)**,  $d_4^{20}$  1.72. It is crystallised twice from distilled water (1ml/g) by adding EtOH, or from warm water (0.5ml/g) by cooling in an ice-salt bath. Dry it in air, then under vacuum. After 3 recrystallisations of ACS grade, it contained Li and B at 0.03 and 0.74 ppm, respectively. It is deliquescent. [Early & Lowry *J Chem Soc* **115** 1387 1919, **121** 963 1922, Hendricks et al. *J Am Chem Soc* **54** 2766 1932.]

**Ammonium perchlorate** [7790-98-9] **M 117.5**,  $d_4^{20}$  1.95,  $pK^{25}$  -2.4 to -3.1 (for HClO<sub>4</sub>). It is recrystallised twice from distilled water (2.5ml/g) between 80° and 0°, and dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. Drying at 110° might lead to slow decomposition to the chloride. **POTENTIALLY EXPLOSIVE.**

**Ammonium peroxydisulfate (ammonium persulfate)** [7727-54-0] **M 228.2**, **m dec when heated wet liberating oxygen**,  $d_4^{20}$  1.98. Recrystallise it at room temperature from EtOH/water. It gradually loses NH<sub>3</sub> on exposure to air. Its solubility is 0.5g/ml at 20°, and 2g/ml at 100°.

**Ammonium reineckate {Reineckate salt, NH<sub>4</sub>[Cr(NH<sub>3</sub>)<sub>2</sub>(SCN)<sub>4</sub>]}** [13573-16-5] **M 336.4 (anhydrous), m 270-273°(dec)**. Crystallise it from water, between 30° and 0°, while working under artificial light. Solutions of reineckate salt (aqueous or alcoholic) decompose slowly at room temperature in the dark (~2 weeks) and more rapidly at higher temperatures or in diffuse sunlight. The solutions are blue in colour and liberate HCN (**POISONOUS**). Store it dry in the dark under a vacuum. [Dakin *Org Synth Coll Vol II* 555 1943.]

**Ammonium selenate** [7783-21-3] **M 179.0**,  $d_4^{20}$  2.19, **m dec on heating**. Crystallise the selenate from water at room temperature by adding EtOH and cooling. Its solubility in H<sub>2</sub>O is 117% at 7° and 197% at 100°. [King *J Phys Chem* **41** 797 1937.]

**Ammonium sulfamate** [7773-06-0] **M 114.1**, **m 132-135°, dec at 160°**. Crystallise it from water at room temperature (1ml/g) by adding EtOH and cooling. [Sisler & Audrieth *Inorg Synth II* 180 1946.]

**Ammonium sulfate** [7783-20-2] **M 132.1**, **m 230°(dec), 280°(dec)**,  $d_4^{20}$  1.77. Crystallise it twice from hot

water containing 0.2% EDTA to remove metal ions, then finally from distilled water. Dry it in a desiccator for 2 weeks over  $\text{Mg}(\text{ClO}_4)_2$ . After 3 recrystallisations, ACS grade had Ti, K, Fe, Na at 11, 4.4, 4.4, 3.2 ppm respectively.

**Ammonium tetrafluoroborate** [13826-83-0] M 104.8,  $\text{pK}^{25}$  2.77 (for  $\text{HBF}_4$ ). Crystallise it from conductivity water (1m/g) between  $100^\circ$  and  $0^\circ$ .

**Ammonium thiocyanate** [1762-95-4] M 76.1, m  $138^\circ(\text{dec})$ ,  $149^\circ(\text{dec})$ ,  $\text{pK}^{25}$  -1.85 (for  $\text{HSCN}$ ), 149. Crystallise it three times from dilute  $\text{HClO}_4$  to give material optically transparent at wavelengths longer than 270nm. It has also been crystallised from absolute MeOH or from acetonitrile.

**Ammonium tungstate (VI)** [11120-25-5] M 283.9,  $\text{pK}_1^{25}$  2.20,  $\text{pK}_2^{25}$  3.70 (for tungstic acid,  $\text{H}_2\text{WO}_4$ ). It crystallises from warm water on adding EtOH and cooling.

**Ammonium (meta) vanadate** [7803-55-6] M 117.0,  $d_{10}^{20}$  2.326. Wash the salt with  $\text{H}_2\text{O}$  until free from  $\text{Cl}^-$  ions and dry it in air. It is soluble in  $\text{H}_2\text{O}$  (5.18g/100ml at  $15^\circ$ , 10.4g/100ml at  $32^\circ$ ) but is more soluble in dilute  $\text{NH}_3$ . It crystallises from conductivity water (20ml/g). When heated at relatively low temperature, it loses  $\text{H}_2\text{O}$  and  $\text{NH}_3$  to give vanadium oxide ( $\text{V}_2\text{O}_5$ ), and at  $210^\circ$  it forms lower oxides. [Baker et al. *Inorg Synth III* 117 1950.] Its solubility in  $\text{H}_2\text{O}$  is 0.52% ( $15^\circ$ ), 1% ( $32^\circ$ ) and 1.6% ( $50^\circ$ ). After washing the technical grade salt with  $\text{H}_2\text{O}$ , it had Na, Mn and U at 0.06, 0.2 and 0.1 ppm, respectively. [Brauer in *Handbook of Preparative Inorganic Chem (Ed. Brauer)* Academic Press Vol II p 1272-1273 1965.]

**Antimony (V) pentafluoride** [7783-70-2] M 216.7, m  $7.0^\circ$ ,  $8.3^\circ$ , b  $141^\circ$ ,  $150^\circ$ ,  $148-150^\circ$ ,  $d_4^{20}$  2.99,  $\text{pK}^{25}$  2.55 [for  $\text{HSb}(\text{OH})_6 = \text{Sb}(\text{OH})_6^- + \text{H}^+$ ]. Purify it by vacuum distillation, preferably in a quartz apparatus, and store it in quartz or aluminum bottles. It is a *hygroscopic* viscous liquid which reacts *violently* with  $\text{H}_2\text{O}$  and is hydrolysed by alkalis. It is **POISONOUS** and attacks the skin. [Woolf & Greenwood *J Chem Soc* 2200 1950, Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 200 1965.]

**Antimony trichloride** [10025-91-9] M 228.1, m  $73^\circ$ , b  $283^\circ$ ,  $\text{pK}_1^{25}$  1.4,  $\text{pK}_2^{25}$  11.0 (11.8),  $\text{pK}_3^{25}$  12.95 (for  $\text{Sb}^{3+}$  aquo). Dry the trichloride over  $\text{P}_2\text{O}_5$  or by mixing it with toluene or xylene and distilling (water is carried off with the organic solvent), then distil it twice under dry nitrogen at 50mm, and sublime it twice in a vacuum into ampoules and seal. It can be crystallised from  $\text{CS}_2$  and is *deliquescent*. It fumes in moist air and is decomposed by  $\text{H}_2\text{O}$  with precipitation of the basic chloride, but forms a clear solution in dilute HCl.

**Antimony trifluoride** [7783-56-4] M 178.8, m  $292^\circ$ , b  $376^\circ$ ,  $d_4^{20}$  4.379. It crystallises from MeOH to remove oxide and oxyfluoride, then it is sublimed under vacuum in an aluminium cup on to a water-cooled copper condenser. Its solubility is 443g/100g in  $\text{H}_2\text{O}$  at  $20^\circ$  and 562g/100g in  $\text{H}_2\text{O}$  at  $30^\circ$  with partial hydrolysis. [Woolf *J Chem Soc* 279 1955, Kwasnik in *Handbook of Preparative Inorganic Chem (Ed. Brauer)* Academic Press Vol I p 199 1963].

**Antimony triiodide** [7790-44-5] M 502.5, m  $167^\circ$ ,  $170^\circ$ , b  $401^\circ$ . It sublimes under vacuum as a red solid with an orange vapour. It hydrolyses to the yellow oxyiodide with  $\text{H}_2\text{O}$ . [Schenk in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 614 1963.]

**Antimony trioxide** [1309-64-4] M 291.5, m  $656^\circ$ . Dissolve the trioxide in the minimum volume of dilute HCl, filter, and add six volumes of water to precipitate the basic antimonous chloride (free from Fe and  $\text{Sb}_2\text{O}_5$ ). The precipitate is redissolved in dilute HCl, and added slowly, with stirring, to a boiling solution (containing a slight excess) of  $\text{Na}_2\text{CO}_3$ . The oxide is filtered off, washed with hot water, then boiled and filtered. The process is repeated until the filtrate gives no test for chloride ions. The product is dried in a vacuum desiccator [Schuhmann *J Am Chem Soc* 46 52 1924]. After one crystallisation (precipitation?), the oxide from a Chinese source had: metal (ppm) Al (8), Ag (0.2), As (56), Cr (6), Ge (0.4), Mn (0.2), Na (16), Ni (2.2) Pb (2.4), Sn (0.4) and V (32). It sublimes in a vacuum at  $400^\circ$ , being yellow on heating and pale buff in colour on cooling. [Schenk in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 615-616 1963.]

**Aqua regia.** This is prepared by slowly adding concentrated  $\text{HNO}_3$  (1 volume) to concentrated hydrochloric acid (3 volumes) in a glass container. This mixture is used to dissolve metals, including noble metals and alloys, as well as minerals and refractory substances. It is done by suspending the material and boiling (**EFFICIENT FUME CUPBOARD — EYE PROTECTION**) to dryness and repeating the process until the residue dissolves in  $\text{H}_2\text{O}$ . If the aqua regia is to be stored for long periods it is advisable to dilute it with one volume of  $\text{H}_2\text{O}$  which will prevent it from releasing chlorine and other chloro and nitrous compounds which are objectionable and toxic. Store it cool in a fume cupboard. However, it is good laboratory practice to prepare it freshly and dispose of it down the fume cupboard sink with copious amounts of water.

**Argon** [7440-37-1] **M 39.95, b -185.6°.** Argon is rendered oxygen-free by passage over reduced copper at  $450^\circ$ , or by bubbling through alkaline pyrogallol and  $\text{H}_2\text{SO}_4$ , then dried with  $\text{CaSO}_4$ ,  $\text{Mg}(\text{ClO}_4)_2$ , or Linde 5A molecular sieves. Other purification steps include passage through Ascarite (**CARE: asbestos** impregnated with sodium hydroxide), through finely divided uranium at about  $800^\circ$  and through a  $-78^\circ$  cold trap.

*Alternatively*, the gas is passed over  $\text{CuO}$  pellets at  $300^\circ$  to remove hydrogen and hydrocarbons, over Ca chips at  $600^\circ$  to remove oxygen and, finally, over titanium chips at  $700^\circ$  to remove nitrogen. It has also been purified by freeze-pump-thaw cycles and by passage over sputtered sodium [Arnold & Smith *J Chem Soc, Faraday Trans 2* 77 861 1981].

**Arsenic acid (arsenic pentoxide hydrate, arsenic V oxide hydrate, orthoarsenic acid)** [12044-50-7] **M 229.8 + xH<sub>2</sub>O, pK<sub>1</sub><sup>25</sup> 2.26, pK<sub>2</sub><sup>25</sup> 6.76, pK<sub>3</sub><sup>25</sup> 11.29 (H<sub>3</sub>AsO<sub>4</sub>).** The acid crystallises from concentrated solutions of boiling conc  $\text{HNO}_3$  as rhombic crystals. Dry it in a vacuum to give the *hemihydrate* (hygroscopic). Heating above  $300^\circ$  yields  $\text{As}_2\text{O}_5$ . [Thaler *Z Anorg Allgem Chem* 246 19 1941, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 601 1963.] **POISONOUS.**

**Arsenic tribromide** [7784-33-0] **M 314.6, m 31.1°, b 89°/11mm, 221°/760mm, d<sub>4</sub><sup>20</sup> 3.67.** Distil it under vacuum. It hydrolyses in  $\text{H}_2\text{O}$ , but less readily than  $\text{AsCl}_3$ . **POISONOUS.** [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 597 1963.]

**Arsenic trichloride (butter of arsenic)** [7784-34-1] **M 181.3, m 31.2°, b 25°/11mm, 130.0°/atm, d<sub>4</sub><sup>20</sup> 2.2.** Reflux the trichloride with arsenic for 4 hours, then fractionally distil it. The middle fraction is stored with sodium wire for two days, then again distilled [Lewis & Sowerby *J Chem Soc* 336 1957]. It fumes in moist air forming the solid hydroxy-chloride [ $\text{AsCl}(\text{OH})_2$ ] and is readily hydrolysed by  $\text{H}_2\text{O}$  to form arsenious acid. **POISONOUS.** [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 596 1963.]

**Arsenic triiodide** [7784-45-4] **M 455.6, m 146°, b 400°/atm, d<sub>4</sub><sup>25</sup> 4.688.** It crystallises from acetone and sublimes below  $100^\circ$ . It is very slowly hydrolysed by  $\text{H}_2\text{O}$  (much more slowly than the chloride). **POISONOUS.** [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 597-598 1963.]

**Arsenic III oxide (arsenic trioxide, arsenious oxide)** [1327-53-3] **M 197.8, three forms: m ~200°(amorphous glass), m 275°(sealed tube, octahedral, common form, sublimes > 125° without fusion but melts under pressure), and m ~312°, pK<sub>1</sub><sup>20</sup> 9.27, pK<sub>2</sub><sup>20</sup> 13.54, pK<sub>3</sub><sup>20</sup> 13.99 (for H<sub>3</sub>AsO<sub>3</sub>).** It crystallises in an *octahedral form* (common form) from  $\text{H}_2\text{O}$  or from dilute  $\text{HCl}$  (1:2), and is then washed, dried and sublimed ( $193^\circ/760\text{mm}$ ). Analytical reagent grade material is suitable for use as an analytical standard after it has been dried at  $105^\circ$  for 1-2 hours or has been left in a desiccator for several hours over conc  $\text{H}_2\text{SO}_4$ .

*Alternatively*,  $\text{As}_2\text{O}_3$  (15g) is dissolved by heating in a mixture of  $\text{H}_2\text{O}$  (60ml) and  $\text{HCl}$  (90g, s.g. 1.1), and crystallisation occurs on cooling, accompanied by brilliant flashes of light [Bandrowski *Z Phys Chem* 17 234 1895].

The *amorphous* form is a colourless transparent glass (**m 200°**) which is obtained when the vapour is slowly condensed below the vaporisation temperature, and should be kept in a sealed tube because it changes to the *octahedral form* (**m 275°**) in the presence of moisture. [Rushton & Daniels *J Am Chem Soc* 48 384 1926.]

A third *monoclinic* form, is obtained by heating the oxide in a sealed tube at  $400^\circ$  (the vitreous, amorphous form

remains at the bottom of the tube) with the *monoclinic* form subliming onto the intermediate part of the tube at 200° (m 312°), and the *octahedral* form deposits at the top of the tube. The transition temperature between the last two forms is ~250°. **POISONOUS (particularly the vapour, handle in a ventilated fume cupboard).** [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 600 1963.]

**Barium (metal)** [7440-39-3] M 137.3, m 727°, b 1640°(1537°)/760mm,  $d_4^{20}$  3.56(3.76). Barium is cleaned by washing with diethyl ether to remove adhering paraffin, then filed in an argon-filled glove box, washed first with ethanol containing 2% conc HCl, then with dry ethanol. It is dried in a vacuum and stored under argon [Addison et al. *J Chem Soc* 3868 1962]. It has also been purified by double distillation under 10mm of argon pressure.

**Barium bromide dihydrate** [7791-28-8] M 333.2, m at 75° loses first H<sub>2</sub>O and at 120° it loses the second H<sub>2</sub>O and melts at 847°. It crystallises from H<sub>2</sub>O (1ml/g) by partial evaporation in a desiccator. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 930 1963.]

**Barium chlorate monohydrate** [10294-38-9 (hydrate), 13477-00-4 (anhydrous)] M 322.3, m 414°. It crystallises from H<sub>2</sub>O (1ml/g) between 100° and 0°, and loses 1H<sub>2</sub>O at 120°. [Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 314 1963.]

**Barium chloride dihydrate** [10326-27-9] M 244.3, m ~120°(dec, hydrate), 963° (anhydrous). It is crystallised twice from water (2ml/g) and dried in an oven to constant weight. The solubilities of the hydrate (% of anhydrous wt) in H<sub>2</sub>O are 31.6 at 0°, 35.7 at 20° and 58.7 at 100°.

**Barium dithionate dihydrate** [13845-17-5] M 333.5, m >150° loses SO<sub>2</sub>, pK<sup>25</sup> 0.49 (for H<sub>2</sub>S<sub>2</sub>O<sub>6</sub>, theory pK<sub>1</sub> -3.4, pK<sub>2</sub> -0.2). It crystallises from water. Its solubility in H<sub>2</sub>O is 7.9% (0°), 15.7% (20°) and 19.9% (30°). [Pfanstiel *Inorg Synth II* 170 1946, Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 397 1963.]

**Barium ferrocyanide hexahydrate** [13821-06-2] M 594.8, m 80°(dec), pK<sub>3</sub><sup>25</sup> 2.57, pK<sub>4</sub><sup>25</sup> 4.35 (for ferrocyanide). It crystallises from hot water (100ml/g).

**Barium fluoride** [7787-32-8] M 175.3, m 1353°, 1368°, b 2260°,  $d_4^{20}$  4.83. Wash it well with distilled H<sub>2</sub>O and dry it in a vacuum. Its solubility in H<sub>2</sub>O is 1.6g (10°), 1.6g (20°) and 1.62g (30°) per L, and is soluble in mineral acids and aqueous NH<sub>4</sub>Cl. It may be stored in glass bottles. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 234 1963.]

**Barium hydroxide octahydrate** [12230-71-6] M 315.5, m 78°, pK<sub>1</sub><sup>25</sup> 13.13, pK<sub>2</sub><sup>25</sup> 13.36. It crystallises from water (1ml/g) and readily absorbs CO<sub>2</sub> from air. It effloresces to the *monohydrate*. It dehydrates to Ba(OH)<sub>2</sub> in dry air at 100°. An aqueous solution (*baryta water*) absorbs CO<sub>2</sub> to form a white precipitate of BaCO<sub>3</sub>.

**Barium hypophosphite monohydrate** [14871-79-5] M 285.4. It precipitates from aqueous solution (3ml/g) on adding EtOH. Its solubility in H<sub>2</sub>O is 28.6% at 17° and 33.3% at 100°. [Klement in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 557 1963.]

**Barium iodate monohydrate** [7787-34-0] M 487.1, m 130°(loses H<sub>2</sub>O), 476°(dec). It crystallises from a large volume of hot water on cooling. Its solubility in H<sub>2</sub>O is 1g/3.35L at 25° and 1g/0.625L at the boiling point. [Lambert & Yasada *Inorg Synth VII* 13 1963.]

**Barium iodide dihydrate** [7787-33-9 (2H<sub>2</sub>O), 13718-50-8 (anhydrous)] M 427.2, m 740°(dec). It crystallises from water (0.5ml/g) by partial evaporation in a desiccator. **POISONOUS.**

**Barium manganate (barium permanganate)** [7787-35-1] M 256.3,  $d_4^{20}$  3.77. Wash the salt with conductivity

H<sub>2</sub>O by decantation until the supernatant gives a faint test for Ba<sup>2+</sup>. Remove excess H<sub>2</sub>O in a vacuum (IMPORTANT), then heat at 100° and the last traces of H<sub>2</sub>O are removed in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. Store it over KOH. It disproportionates in hot H<sub>2</sub>O or dilute acid into Ba(MnO<sub>2</sub>)<sub>2</sub> and MnO<sub>2</sub>, and is a mild oxidant. [Schlezinger & Siems *J Am Chem Soc* **46** 1965 1924, Nyholm & Woolliams *Inorg Synth XI* 56 1968.]

**Barium nitrate** [10022-31-8] **M 261.4, m 593°(dec)**. Crystallise it twice from water (4ml/g) and dry it overnight at 110°. It decomposes at higher temperatures to give mostly the oxide and the peroxide with only a little of the nitrite. **POISONOUS**. [Ehrlich in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 941 1963.]

**Barium nitrite monohydrate** [7787-38-4] **M 247.4, m 217°(dec)**. Barium nitrite crystallises from water (1ml/g) on cooling in an ice-salt bath. **POISONOUS**.

**Barium perchlorate** [13465-95-7] **M 336.2, m 505°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Recrystallise the perchlorate twice from water. [Schmeisser in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 320 1963.]

**Barium sulfate** [7722-43-7] **M 233.4, m >1580°**. Wash the sulfate five times by decantation with hot distilled water, dialyse it against distilled water for one week, then freeze-dry and dry in an oven at 105° to constant weight (~12 hours). It is almost insoluble in H<sub>2</sub>O (its solubility is 0.0024g/L at 25°).

**Barium tetrathionate** [82203-66-5] **M 361.6**. Purify the tetrathionate by dissolving it in a small volume of water and precipitating it with EtOH below 5°. After drying, the salt is stored in the dark at 0°. [see Fréhér in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 399-401 1963 for potassium tetrathionate K<sub>2</sub>S<sub>4</sub>O<sub>4</sub>.]

**Barium thiocyanate dihydrate** [2092-17-3] **M 289.6, pK<sup>25</sup> -1.85 (for HSCN)**. It is crystallised from water (2.5ml/g) by partial evaporation in a desiccator. It is deliquescent. [Herstein *Inorg Synth III* 24 1950.]

**Barium thiosulfate** [35112-53-9] **M 249.5, m 220°(dec), pK<sub>1</sub><sup>25</sup> 0.6, pK<sub>2</sub><sup>25</sup> 1.74 (for H<sub>2</sub>S<sub>2</sub>O<sub>3</sub>)**. It is very slightly soluble in water and is washed repeatedly with chilled water and dried in air at 40°.

**Beryllium carbonate (BeCO<sub>3</sub>)** [744998-97-8] **M 69.0**. The commercial carbonate contains more than 1% of impurities. It is best purified by converting it to the basic acetate (see [1332-52-1]) which should be sublimed, converted into the nitrate with HNO<sub>3</sub> see below [13587-99-4]) and evaporated in a Pt dish until free from excess acid. The residual nitrate salt is dissolved in a small volume of H<sub>2</sub>O and enough ammonium carbonate solution is added to redissolve the beryllium carbonate that precipitates and gives a clear solution. This is then evaporated in a Pt dish until course BeCO<sub>3</sub> separates. It is washed thoroughly with pure H<sub>2</sub>O, then distilled EtOH and dried *in vacuo*. Be(OH)<sub>2</sub>, BeO and Be salts are generally insoluble in NH<sub>3</sub> but dissolve in **aqueous** (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub> solution to form the carbonate. [Ehrlich in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 893 1963.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium chloride (BeCl<sub>2</sub>)** [7787-47-5] **M 79.9, m 399°, 405°, b 488°/atm, d<sup>25</sup> 1.89**. The anhydrous chloride is obtained by heating Be powder or chips (or BeO + C) with Cl<sub>2</sub> gas in a furnace at 300-350° for *ca* 40 hours when anhydrous BeCl<sub>2</sub> sublimes out as a white to faintly yellow crystalline mass or orthorhombic crystals. It is purified further by sublimation at ~300° *in vacuo*. It is very hygroscopic and reacts exothermically with H<sub>2</sub>O which becomes acidic. The solid should be stored in a tightly stoppered vessel and handled in a dry atmosphere much like anhydrous AlCl<sub>3</sub> and has similar catalytic properties to it. It is insoluble in \*C<sub>6</sub>H<sub>6</sub> and toluene but is soluble EtOH, Et<sub>2</sub>O, pyridine and CS<sub>2</sub> with which it probably complexes. The conductivity of fused BeCl<sub>2</sub> is ~0.001 of that of fused NaCl implying that 1 molecule in 1000 is ionised. [Tannenbaum *Inorg Synth* **5** 22 1957, Ehrlich in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 889 1963.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium hydroxide [Be(OH)<sub>2</sub>]** [13327-32-7] **M 43.0, d<sup>25</sup> 1.89**. The hydroxide exists in two forms, a meta-

stable  $\alpha$ -form and a stable  $\beta$ -form. The former is obtained in amorphous form by precipitating a Be salt with  $\text{NH}_3$  in the absence of  $\text{CO}_2$  and allowing it to age by prolonged heating (~24 hours) with  $\text{NH}_4\text{OH}$  solution.  $\beta$ - $\text{Be}(\text{OH})_2$  is made by boiling a saturated solution of the amorphous  $\alpha$ -form in 10N aqueous  $\text{NaOH}$  until a permanent turbidity is attained. On slow cooling fine crystals (regular double pyramids) of the  $\beta$ -form separate, and are purified by washing with warm  $\text{H}_2\text{O}$  until the washings are no longer alkaline, then drying at  $80^\circ$ . The mother liquors can be reused in the above process. It is very slightly soluble in  $\text{H}_2\text{O}$  and dilute  $\text{NaOH}$ , but both forms are amphoteric and are soluble in hot concentrated aqueous  $\text{NaOH}$  solutions and in acids. It dehydrates on strong heating to  $\text{BeO}$ . Several sodium beryllates are known. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 894, 895 1963.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium nitrate** [ $\text{Be}(\text{NO}_3)_2$ ] [13587-99-4;  $3\text{H}_2\text{O}$  7787-55-5] M 133.0, m  $\sim 60^\circ$  (for  $3\text{H}_2\text{O}$ ). The nitrate does not crystallise easily and is very deliquescent. It can be prepared from the metal by digesting it with conc  $\text{HCl}$  until most of it has dissolved, filter any insoluble material through a 0.45 micron acetate filter. Then at low boil, add conc  $\text{HNO}_3$  with low heat to reduce to a slurry. Repeat with careful addition of  $\text{HNO}_3$  as many times as necessary (3-6 additions) until the slurry tests negative for chloride ions with 0.1N aqueous  $\text{AgNO}_3$ . Then isolate the nitrate as described below [J. Papa personal communication, 2011, see Chapter 1 p 38]. Alternatively, it is obtained by heating recrystallised, then sublimed, basic beryllium acetate  $\{\text{Be}_4\text{O}(\text{OAc})_6$ , see [1332-52-1]\}, with pure concentrated  $\text{HNO}_3$  in a degassed pyrex container until the volatiles are removed. Heat well below  $1100^\circ$ , so as not to decompose it to give pure  $\text{BeO}$ . It picks up moisture to give the white to pale yellow *trihydrate*. Store it in a well stoppered bottle in a cool place. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 893 1963.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium oxide** [1304-56-9] M 25.0, m  $2530^\circ$ , d<sup>25</sup> 3.01. It is prepared by calcining the nitrate at  $1100^\circ$  {see  $\text{Be}(\text{NO}_3)_2$  above [13587-99-4]}, and this oxide always contains  $\sim 0.35\text{ml}$  of gas ( $\text{N}_2$ ,  $\text{O}_2$ ) per gram of oxide. Alternatively, freshly prepared beryllium carbonate is calcined in a Pt boat placed in an electric furnace at  $900^\circ$ .  $\text{BeO}$  is amphoteric, dissolving slowly in acids and in alkalis but if heated strongly it becomes refractory and difficult to dissolve except in hydrofluoric acid. Like ceramics it is a good electrical insulator, but it is a good conductor of heat like some metals. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 893 1963.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium potassium fluoride** [7787-50-0] M 105.1, m  $\sim 350^\circ$ . It crystallises from hot water (25ml/g). Beryllium compounds are potentially **carcinogenic**.

**Beryllium sulfate tetrahydrate** [7787-56-6:] M 177.1, m  $\sim 100^\circ(\text{dec})$ , d<sup>25</sup> 1.713, pK<sub>1</sub><sup>25</sup> 3.2, pK<sub>2</sub><sup>25</sup>  $\sim 6.5$  ( $\text{Be}^{2+}$ ). It crystallises from a concentrated solution of weak aqueous  $\text{H}_2\text{SO}_4$  (with partial evaporation), and is dried in air. It is quite soluble in  $\text{H}_2\text{O}$ . The *tetrahydrate* loses  $\text{H}_2\text{O}$  at  $\sim 100^\circ$  to give the *dihydrate*. [cf Gmelin's *Beryllium* (8<sup>th</sup> edn) 26 169 1930.] Beryllium compounds are potentially **carcinogenic**.

**Bismuth** [7440-69-9] M 209.0, m  $271\text{-}273^\circ$ , b  $1450^\circ$ . Melt it in an atmosphere of dry helium, then filter it through dry Pyrex wool to remove any bismuth oxide present [Mayer et al. *J Phys Chem* 64 238 1960].

**Bismuth trichloride** [7787-60-2] M 315.3, m  $233.6^\circ$ , pK<sup>25</sup> 1.58 for hydrolysis ( $\text{Bi}_3^- = \text{BiOH}_2^+ + \text{H}^+$ ). Sublime the trichloride under high vacuum, or dry it under a current of  $\text{HCl}$  gas, followed by fractional distillation, once under  $\text{HCl}$  and once under argon. It is *deliquescent*. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 621 1963.]

**Boric acid (orthoboric acid, boracic acid)** [10043-35-3] M 61.8, m  $171^\circ$ , pK<sup>25</sup> 9.23. Crystallise the acid three times from  $\text{H}_2\text{O}$  (3ml/g) between  $100^\circ$  and  $0^\circ$ , after filtering through sintered glass. Dry it to constant weight over metaboric acid in a desiccator. It is steam volatile. After two recrystallisations of ACS grade, it had Ag at 0.2 ppm. Its solubility (%) in  $\text{H}_2\text{O}$  is 2.66 at  $0^\circ$ , 4.0 at  $12^\circ$  and 24 at  $80^\circ$ . At  $100^\circ$  it loses  $\text{H}_2\text{O}$  to form *metaboric acid* ( $\text{HBO}_2$ ). When it is heated to redness or slowly to  $200^\circ$ , or over  $\text{P}_2\text{O}_5$  *in vacuo*, it dehydrates to *boric anhydride* ( $\text{B}_2\text{O}_3$ ) [1303-82-6] to give a white hard glass or crystals with m  $\sim 294^\circ$ . The glass softens on

heating and liquefies at red heat. It is an astringent, a fungicide and an antibacterial. [McCulloch *J Am Chem Soc* **59** 2650 1937, Kelly *J Am Chem Soc* **63** 1137 1941, Taylor & Cole *J Chem Soc* 70 1926, Conti *J Soc Chem Ind* **44** 343T 1925.]

**Boron trichloride (trichloroborane)** [10294-34-5] **M 117.2, m -107°, b 0°/476mm, 12.5°/atm, d<sup>0</sup> 1.3728, d<sub>4</sub><sup>12</sup> 1.35** (also available as a 1.0M solution in toluene, *p*-xylene, CH<sub>2</sub>Cl<sub>2</sub>, heptane or hexanes). Purify it (from chlorine) by passage through two mercury-filled bubblers, then fractionally distil it under a slight vacuum. In a more extensive purification the nitrobenzene addition compound is formed by passage of the gas over nitrobenzene in a vacuum system at 10°. Volatile impurities are removed from the crystalline yellow solid formed by pumping at -20°, and the BCl<sub>3</sub> is recovered by warming the addition compound at 50°. Passage through a trap at -78° removes entrained nitrobenzene, the BCl<sub>3</sub> finally condensing in a trap at -112° [Brown & Holmes *J Am Chem Soc* **78** 2173 1956]. Alternatively, purify it by condensing it into a trap cooled in acetone/Dry-ice, where it is pumped for 15 minutes to remove volatile impurities. It is then warmed, recondensed and again pumped. [Gamble *Inorg Synth III* 27 1950.]

It fumes in moist air. **TOXIC, do not breath in the vapours.** It forms addition compounds with Et<sub>2</sub>O; and with dimethyl sulfide it forms Me<sub>2</sub>S·BCl<sub>3</sub> [5523-19-3] **M 179.3**, which has **m 88-90°**, and is a convenient solid form of it (also commercially available as a 2.0M solution in CH<sub>2</sub>Cl<sub>2</sub>) — **beware** of the foul odour of Me<sub>2</sub>S; work in an efficient fume cupboard and absorb the sulfide vapours by flushing them with a stream of N<sub>2</sub> into lead acetate solution which will precipitate insoluble black PbS. [Gamble *Inorg Synth 3* 27 1950, Gerrard & Leppert *Chem Rev* **58** 1081 1958.]

**Boron trifluoride** [7637-07-2] **M 67.8, b -101°/760mm.** The usual impurities-bromine, BF<sub>5</sub>, HF and non-volatile fluorides-are readily separated by distillation. Brown and Johannesen [*J Am Chem Soc* **72** 2934 1950] passed BF<sub>3</sub> into benzonitrile at 0° until the latter was saturated. Evacuation to 10<sup>-5</sup>mm then removed all traces of SiF<sub>4</sub> and other gaseous impurities. [A small amount of the BF<sub>3</sub>-benzonitrile addition compound sublimes and is collected in a U-tube cooled to -80°]. The pressure is raised to 20mm by admitting dry air, and the flask containing the BF<sub>3</sub> addition compound is warmed with hot water. The BF<sub>3</sub> that evolves is passed through a -80° trap (to condense any benzonitrile) into a tube cooled in liquid air. The addition compound with anisole can also be used. BF<sub>3</sub> can be dried by passing it through H<sub>2</sub>SO<sub>4</sub> saturated with boric oxide. It fumes in moist air. [It is commercially available as a 1.3M solution in MeOH or PrOH.] [Booth & Wilson *Inorg Synth I* 21 1939, Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** pp 219-222 1963.] **TOXIC.**

**Bromine** [7726-95-6] **M 159.8, m -7.2°, b 59°, d<sub>4</sub><sup>20</sup> 3.102, n<sub>D</sub><sup>20</sup> 1.661.** Reflux the brown liquid with solid KBr and distil; dry the distillate by shaking it with an equal volume of conc H<sub>2</sub>SO<sub>4</sub>, then redistil it. The H<sub>2</sub>SO<sub>4</sub> treatment can be replaced by direct distillation from BaO or P<sub>2</sub>O<sub>5</sub>. A more extensive purification [Hildenbrand et al. *J Am Chem Soc* **80** 4129 1958] is to reflux about 1L of bromine for 1 hour with a mixture of 16g of CrO<sub>3</sub> in 200ml of conc H<sub>2</sub>SO<sub>4</sub> (to remove organic material). The bromine is distilled into a clean, dry, glass-stoppered bottle, and chlorine is removed by dissolving *ca* 25g of freshly fused CsBr in 500ml of the bromine and standing overnight. To remove HBr and water, the bromine is then distilled back and forth through a train containing alternate tubes of MgO and P<sub>2</sub>O<sub>5</sub>. [Schmeisser in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 275 1963.] **HIGHLY TOXIC.**

**Bromine pentafluoride** [7789-30-2] **M 174.9, m -60.5°, b 41.3°, d<sub>25</sub><sup>25</sup> 2.466.** Purify it *via* its KF complex, as described for chlorine trifluoride. [Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** pp 158-159 1963.] **HIGHLY TOXIC.**

**Cadmium** [7440-43-9] **M 112.4, m 321.1°, b 767°, d<sub>4</sub><sup>20</sup> 8.642.** Any oxide contaminant is removed by filtering the molten metal, under vacuum, through quartz wool. Its solubility in Hg is 5.2% (18°), and it is soluble in mineral acids. [Wagenknecht & Juza in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **II** p 1092 1965.]

**Cadmium bromide** [7789-42-6 (anhydrous), 13464-92-1 (4H<sub>2</sub>O)] **M 272.2, m 566°, b 863°, 963°, d<sub>4</sub><sup>20</sup> 5.192.**

Crystallise it from water (0.57g/ml at 10°, 1.26g/ml at 100°) between and 0°, and dry it at 110°. It forms the *monohydrate* below 36° and the *tetrahydrate* above 36°. It is hygroscopic. [Wagenknecht & Juza in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1096 1965.]

**Cadmium chloride** [10108-64-2] **M 183.3, m 568°, b 960°, d<sub>4</sub><sup>20</sup> 4.06.** Crystallise it from water (1ml/g) by addition of EtOH and cooling. [Pray *Inorg Synth* V 153 1957, Wagenknecht & Juza in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1093 1965.]

**Cadmium fluoride** [7790-79-6] **M 150.4, m >1000°, b 1748°, d<sub>4</sub><sup>20</sup> 6.35.** Crystallise it by dissolving it in hot water (25ml/g at room temperature) at 60°, filtering, then cooling. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 243 1963.]

**Cadmium iodide** [7790-80-9] **M 366.2, m 388°, b 787°, d<sub>4</sub><sup>20</sup> 5.66.** Crystallise it from ethanol (2ml/g) by partial evaporation. [Wagenknecht & Juza in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1096 1965.]

**Cadmium nitrate tetrahydrate** [10022-68-1] **M 308.5, m 59.5°.** Crystallise the nitrate from water (0.5ml/g) by cooling in ice-salt. [Gmelin's *Cadmium* (8th edn) 33 pp 76-78 1925, Suppl p 446 1959.]

**Cadmium potassium iodide** [13601-63-3] **M 532.2.** Crystallise it from ethanol by partial evaporation.

**Cadmium sulfate** [7790-84-3 (for 3CdSO<sub>4</sub> 8H<sub>2</sub>O), 10124-36-4 (anhydrous)] **M 208.4 (anhydrous), 769.5 (hydrate).** The sulfate crystallises from distilled water as a *hydrate* by partial evaporation in a desiccator. It gives the *monohydrate* on heating at 80°. It is insoluble in EtOH, Me<sub>2</sub>CO or EtOAc. It forms a white precipitate of Cd(OH)<sub>2</sub> with aqueous NH<sub>3</sub> which dissolves in excess of NH<sub>3</sub> to form soluble [Cd(NH<sub>3</sub>)<sub>4</sub>]SO<sub>4</sub>. [Gmelin's *Cadmium* (8th edn) 33 p 121 1925, Suppl pp 609-610 1959.]

**Calcium** [7440-70-2] **M 40.1, m 845°.** Clean the metal by washing it with ether to remove adhering paraffin, file the surface in an argon-filled glove box, and wash it with ethanol containing 2% of conc HCl. Then wash it with dry ethanol, dry it in a vacuum and store it under pure argon [Addison et al. *J Chem Soc* 3868 1962].

**Calcium bromide monohydrate** [62648-72-0, 71626-99-8 (xH<sub>2</sub>O), 7789-41-5 (anhydrous)] **M 217.9, d<sub>4</sub><sup>20</sup> 3.35.** Crystallise the bromide from EtOH or Me<sub>2</sub>CO. It loses H<sub>2</sub>O on heating, is anhydrous at 750°, then it loses Br at higher temperatures. It is *deliquescent*. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 930 1963.]

**Calcium chloride (anhydrous)** [10043-52-4] **M 111.0, m 772°, b >1600°, d<sub>4</sub><sup>15</sup> 2.15.** It is available as fused granules or cubic crystals. It is very *hygroscopic*, very soluble in H<sub>2</sub>O (exothermic), and EtOH. Store it in a tightly closed container. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 931 1963.]

**Calcium chloride dihydrate** [10035-04-8] **M 147.0, m 175°(dehydrates), 772°(dec).** Crystallise it from ethanol. It is *hygroscopic*. It loses H<sub>2</sub>O at 200° so it can be dried at high temperatures to dehydrate it. The *hexahydrate* [7774-34-7] has **m 30°** and **d 1.67**.

**Calcium dithionite** [13812-88-9] **M 168.2, m dec on heating.** Crystallise it from water, or water followed\* by acetone and dry it in air at room temperature.

**Calcium hexacyanoferrate (II) undecahydrate** [13821-08-4] **M 490.3.** Recrystallise it three times from conductivity H<sub>2</sub>O and dry it in air to constant weight over the partially dehydrated salt. [James *Trans Faraday Soc* 45 855 1949.] Alternatively, the Ca salt can be purified by precipitation with absolute EtOH in the cold (to avoid oxidation) from an air-free saturated aqueous solution. The pure lemon yellow crystals are centrifuged, dried in a vacuum desiccator first over dry charcoal for 24 hours, then over partly dehydrated salt and stored in a dark glass stoppered bottle. No deterioration occurs after 18 months. No trace of Na, K or NH<sub>4</sub> ions can be

detected in the salt from the residue after decomposition of the salt with conc  $\text{H}_2\text{SO}_4$ . Analyses indicate 11 mols of  $\text{H}_2\text{O}$  per mol of salt. The solubility in  $\text{H}_2\text{O}$  is 36.45g (24.9°) and 64.7g (44.7°) per 100g of solution. [Farrow *J Chem Soc* 50 1926.]

**Calcium hydroxide** [1305-62-0] **M 74.1, m loses  $\text{H}_2\text{O}$  on heating,  $\text{pK}^{25}$  12.7 (for  $\text{Ca}^{2+}$ ).** Heat analytical grade calcium carbonate at  $1000^\circ$  during 1 hour. Allow the resulting oxide to cool and add slowly to water. Heat the suspension to boiling, cool and filter through a sintered glass funnel of medium porosity (to remove soluble alkaline impurities). Dry the solid at  $110^\circ$  and crush it to a uniformly fine powder. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 934 1963.]

**Calcium iodate** [7789-80-2 ( $\text{H}_2\text{O}$ )] **M 389.9, m  $>540^\circ$ ,  $\text{pK}^{25}$  0.79 (for  $\text{HIO}_3$ ).** Crystallise it from water (100ml/g at  $100^\circ$  and 100ml/0.1g at  $0^\circ$ ). [Bahl & Singh *J Indian Chem Soc* 17 397 1940.]

**Calcium iodide x-hydrate** [71626-98-7 ( $x\text{H}_2\text{O}$ ), 10102-68-8 (anhydrous)] **M 293.9 (for  $4\text{H}_2\text{O}$ ), m  $740^\circ$ , b  $1100^\circ$ .** Dissolve the salt in acetone, which is then diluted and evaporated. This drying process is repeated twice, then the  $\text{CaI}_2$  is recrystallised from acetone/diethyl ether and stored over  $\text{P}_2\text{O}_5$ . It is very *hygroscopic* when anhydrous and is light sensitive [Cremlyn et al. *J Chem Soc* 528 1958]. The yellow *hexahydrate* has **m  $42^\circ$** . It is soluble in  $\text{H}_2\text{O}$ , MeOH, EtOH and  $\text{Me}_2\text{CO}$  but insoluble in  $\text{Et}_2\text{O}$ .

**Calcium nitrate tetrahydrate** [13477-34-4] **M 236.1, m  $45^\circ$ (hydrate),  $560^\circ$ (anhydrous).** Crystallise the nitrate four times from water (0.4ml/g) by cooling in a  $\text{CaCl}_2$ -ice freezing mixture. The *tetrahydrate* is dried over conc  $\text{H}_2\text{SO}_4$  and stored over  $\text{P}_2\text{O}_5$ , to give the *anhydrous salt*. It is *deliquescent*. After 3 recrystallisations of ACS grade salt, it had Co, Fe, Mg, Sr and Zn at 0.2, 1.0, 0.02, 10 and 0.02 ppm resp. [Bassett & Taylor *J Chem Soc* 105 1926 1914.]

**Calcium nitrite dihydrate** [13780-06-8 (30% w/w aqueous solution)] **M 150.1(hydrate), m dec on heating,  $d_4^{20}$  2.22.** Crystallise it from hot water (1.4ml/g) by adding ethanol and cooling to give the hydrate. It is *deliquescent*. [Ray & Ogg *J Am Chem Soc* 79 265 1957.]

**Calcium permanganate tetrahydrate** [10118-76-0 (anhydrous)] **M 350.0 (for  $4\text{H}_2\text{O}$ ).** Crystallise it from water (3.3ml/g) by partial evaporation in a desiccator. It is *deliquescent*. Note that it loses oxygen more readily than the potassium salt. It is an oxidising agent.

**Calcium sulfate dihydrate** [10101-41-4] **M 172.1, m 150(dec),  $d_4^{20}$  2.32.** It loses only part of its  $\text{H}_2\text{O}$  at  $100$ - $150^\circ$  (see below). It is soluble in  $\text{H}_2\text{O}$  and very slowly soluble in glycerol. It is insoluble in most organic solvents.

**Calcium sulfate hemihydrate** [10034-76-1] **M 145.2.** Its solubility in  $\text{H}_2\text{O}$  is 0.2parts/100 at  $18.75^\circ$ . It dehydrates completely  $>650^\circ$ . Dry it below  $300^\circ$  to give a solid with estimated pore size *ca* 38% of volume. *Anhydrous*  $\text{CaSO}_4$  (**Drierite**) has a high affinity for  $\text{H}_2\text{O}$  and will absorb 6.6% of its weight of  $\text{H}_2\text{O}$  to form the *hemihydrate* (gypsum). It sets to a hard mass with  $\text{H}_2\text{O}$ ; hence it should be kept in a tightly sealed container. The solubility of gypsum in  $\text{H}_2\text{O}$  is unusual: 0.176% at  $0^\circ$ , 0.209% at  $30^\circ$ , 0.210 at  $40^\circ$ , 0.204 at  $50^\circ$  and 0.200 at  $60^\circ$ . [Hulett *J Am Chem Soc* 27 49 1905, James & Partington *J Chem Soc* 107 1019 1915, Namba *J Soc Chem Ind* 40 2797 1920.]

**Calcium thiosulfate** [10124-41-1] **M 152.2, m  $43$ - $49^\circ$ ,  $\text{pK}_1^{25}$  0.6,  $\text{pK}_2^{25}$  1.74 (for  $\text{H}_2\text{S}_2\text{O}_3$ ).** Recrystallise the thiosulfate from water below  $60^\circ$  in a  $\text{N}_2$  atmosphere, followed by drying with EtOH and  $\text{Et}_2\text{O}$ . Store it in a refrigerator. The *hexahydrate* can decompose spontaneously at  $43$ - $49^\circ$ . Store it in a cool closed container. [Pethybridge & Taba *J Chem Soc, Faraday Trans 1* 78 1331 1982.]

**Carbon dioxide** [124-38-9] **M 44.0, sublimes at  $-78.5^\circ$ ,  $\text{pK}_1^{25}$  6.35,  $\text{pK}_2^{25}$  10.33 (for  $\text{H}_2\text{CO}_3$ ).** Pass the gas over CuO wire at  $800^\circ$  to oxidise CO and other reducing impurities (such as  $\text{H}_2$ ), then over copper dispersed on Kieselguhr at  $180^\circ$  to remove  $\text{O}_2$ . Drying it at  $-78^\circ$  removes the water vapour. Final purification is by vacuum

distillation at liquid nitrogen temperature to remove non-condensable gases [Anderson et al. *J Chem Soc* 3498 1962]. Sulfur dioxide contaminant can be removed at 450° using silver wool combined with a plug of platinised quartz wool. Halogens are removed by using Mg, Zn or Cu, heated to 450°. [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 647 1963.]

**Carbon disulfide** [75-15-0] **M 76.1, b 46.3°, d<sub>4</sub><sup>20</sup> 1.264, n<sub>D</sub><sup>20</sup> 1.627.** Shake it for 3 hours with three portions of KMnO<sub>4</sub> solution (5g/L), twice for 6 hours with mercury (to remove sulfide impurities) until no further darkening of the interface occurs, and finally with a solution of HgSO<sub>4</sub> (2.5g/L) or cold, saturated HgCl<sub>2</sub>. Dry it with CaCl<sub>2</sub>, MgSO<sub>4</sub>, or CaH<sub>2</sub> (with further drying by refluxing over P<sub>2</sub>O<sub>5</sub>), followed by fractional distillation in diffuse light. **Alkali metals cannot be used as drying agents.** It has also been purified by standing with bromine (0.5ml/L) for 3-4 hours, shaking rapidly with KOH solution, then copper turnings (to remove unreacted bromine), and drying with CaCl<sub>2</sub>. CS<sub>2</sub> is highly **TOXIC** has a **foul odour** and is highly **FLAMMABLE**. *Work in a good fumehood.*

Small quantities of CS<sub>2</sub> have been purified (including removal of hydrocarbons) by mechanical agitation of a 45-50g sample with a solution of 130g of sodium sulfide in 150ml of H<sub>2</sub>O for 24 hours at 35-40°. The aqueous sodium thiocarbonate solution is separated from unreacted CS<sub>2</sub>, then precipitated with 140g of copper sulfate in 350g of water, with cooling. After filtering off the copper thiocarbonate, it is decomposed by passing steam into it. The distillate is separated from H<sub>2</sub>O and distilled from P<sub>2</sub>O<sub>5</sub>. [Ruff & Golla *Z Anorg Chem* **138** 17 1924, *Beilstein* **3** IV 395.]

**Carbon monoxide** [630-08-0] **M 28.0, m -200°, b -191.5°.** Iron carbonyl is a likely impurity in CO stored under pressure in steel tanks. It can be decomposed by passing the gas through a hot porcelain tube at 350-400°. Passage through alkaline pyrogallol solution removes oxygen (and CO<sub>2</sub>). Removal of CO<sub>2</sub> and water are effected by passage through soda-lime followed by Mg(ClO<sub>4</sub>)<sub>2</sub> or P<sub>2</sub>O<sub>5</sub> and collected over Hg. Carbon monoxide can be condensed and distilled at -195°. It is sparingly soluble in H<sub>2</sub>O but is readily absorbed by a solution of CuCl in HCl to give the white crystalline adduct CuCl.CO.2H<sub>2</sub>O. It burns in air with a bright blue flame but a mixture of 2 volumes of CO and 1 volume of O<sub>2</sub> explode when kindled, although in a small jar the combustion is not violent. **HIGHLY POISONOUS gas** as it reacts with haemoglobin to form bright red carboxyhaemoglobin which is stable and not readily decomposed by oxygen. The **TOXIC LEVEL** of CO is 50ppm (~55mg/m<sup>3</sup>). **The ANTIDOTE should be at hand and always available in laboratories using CO; and staff should be trained to administer it.** [Gilliland & Blanchard *Inorg Synth* **II** 81 1946, Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 645-646 1963.]

**Carbonyl bromide** [593-95-3] **M 187.8, b 64.5°/760mm.** Purify it by distillation from Hg and from powdered Sb to remove free bromine, then distil it in a slight vacuum to remove volatile SO<sub>2</sub> (the major impurity) [Carpenter et al. *J Chem Soc, Faraday Trans* **2** 384 1977]. **TOXIC.**

**Carbonyl sulfide (COS)** [463-58-1] **M 60.1, m -138°, b -47.5°, -50°.** Purify the gas by scrubbing it through three consecutive fritted washing flasks containing conc NaOH at 0° (to remove HCN), and then through conc H<sub>2</sub>SO<sub>4</sub> (to remove CS<sub>2</sub>) followed by a mixture of NaN<sub>3</sub> and NaOH solution; or passed through traps containing saturated aqueous lead acetate, then through a column of anhydrous CaSO<sub>4</sub>. Then it is freeze-pumped repeatedly and distilled through a trap packed with glass wool and cooled to -130° (using an *n*-pentane slurry). It liquefies at 0°/12.5mm. Use stainless steel containers. The gas is stored over conc H<sub>2</sub>SO<sub>4</sub>. [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 654 1963.] **TOXIC.**

**Ceric ammonium nitrate** [16774-21-3] **M 548.2, pK<sub>1</sub><sup>25</sup> -1.15, pK<sub>2</sub><sup>25</sup> -0.72, pK<sub>3</sub><sup>25</sup> 1.68, pK<sub>4</sub><sup>25</sup> 2.29 (for aquo Ce<sup>4+</sup>).** Ceric ammonium nitrate (125g) is warmed with 100ml of dilute HNO<sub>3</sub> (1:3 v/v) and 40g of NH<sub>4</sub>NO<sub>3</sub> until it dissolves, and filtered through a sintered-glass funnel. The solid which separates on cooling in ice is filtered off on a sintered funnel (at the pump) and air is sucked through the solid for 1-2 hours to remove most of the nitric acid. Finally, the solid is dried at 80-85°.

**Cerium(III) sulfate** [cerous sulfate, Ce<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>] [13454-94-9 anhydrous; 10450-59-6 8H<sub>2</sub>O; 13550-47-5 12 H<sub>2</sub>O] **M 568.4 (anhydr), d<sub>4</sub><sup>25</sup> 2.89, pK<sup>25</sup> 9.29 (for hydrolysis of Ce<sup>3+</sup>).** To prepare cerous sulfate, cerium

oxide [CeO<sub>2</sub>, 0.3g] or chloride is dissolved in hot 6N H<sub>2</sub>SO<sub>4</sub> (20ml), filtered through a sintered glass funnel and allowed to crystallise over concentrated H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. The crystals are filtered off, washed twice with H<sub>2</sub>O (10ml) and once with EtOH (10ml), and dried in air for 4 hours to provide the *pentahydrate*. Alternatively, a neutral or slightly acidic (with H<sub>2</sub>SO<sub>4</sub>) aqueous solution of the sulfate is treated with ¾ of its volume of EtOH. The salt is thus obtained rapidly and quantitatively without the evaporation stage. The *anhydrous* sulfate is formed by heating any sulfate hydrate salt at 400-500°, but is hygroscopic. The hydrate loses most of its H<sub>2</sub>O at ca 250°, and above 650° it loses SO<sub>3</sub> also to give the basic salt. [Wendlandt *J Inorg Nucl Chem* 7 51 1958, Wetzell in *Handbook of Preparative Inorganic Chem (Ed. Brauer)* Academic Press Vol II p 1156 1965.] When a fairly concentrated solution of the anhydrous salt in ice H<sub>2</sub>O, is allowed to evaporate at 40-50°, the *octahydrate* salt separates as colourless rhombic pyramids [Koppel *Z Anorg Chem* 41 385 1904]. The *dodecahydrate* is obtained in fine needles when a clear fairly concentrated solution of cerous sulfate is allowed to evaporate over concentrated H<sub>2</sub>SO<sub>4</sub> in a desiccator in a refrigerator [Koppel *Z Anorg Chem* 41 385 1904]. [Vanino *Handbuch der Präparativen Chemie* Vol I, 3<sup>rd</sup> Edn, pp 754-755, Stuttgart 1925.]

**Cerium(IV) sulfate** [ceric sulfate, Ce(SO<sub>4</sub>)<sub>2</sub>] [13590-82-4 *anhydrous*] M 332.2, d<sub>4</sub><sup>25</sup> 3.01, pK<sub>1</sub><sup>25</sup>-1.15, pK<sub>2</sub><sup>25</sup>-0.72, pK<sub>3</sub><sup>25</sup>1.68, pK<sub>4</sub><sup>25</sup>2.29 (for hydrolysis of Ce<sup>4+</sup>). Ceric sulfate is prepared by heating pure ceric oxide (CeO<sub>2</sub>) with an excess of concentrated H<sub>2</sub>SO<sub>4</sub> for an hour, cooling, adding glacial acetic acid, stirring, allowing to settle and decanting off. The process is repeated several times and the yellow orthorhombic crystals of the sulfate are filtered off (sintered glass), washed with glacial acetic acid and dried in a vacuum desiccator over NaOH/KOH and/or soda lime. [Vanino *Handbuch der Präparativen Chemie* Vol I, 3<sup>rd</sup> Edn, pp 754-755, Stuttgart 1925, Meyer & Aufrecht *Chem Ber* 37 140 1904.] It is a strong oxidising agent (oxidation potential of a 1 to 8N H<sub>2</sub>SO<sub>4</sub> solution at 25° is 1.42 ±0.1 volts), and like ceric ammonium sulfate is used in volumetric analysis. Solutions in dilute H<sub>2</sub>SO<sub>4</sub> are standardised by titration with pure arsenious oxide, iron or ferrous ammonium sulfate. Although it is its own end-point indicator (Ce<sup>4+</sup> ions are yellow and Ce<sup>3+</sup> ions are colourless), the end point is not as sharp as that of permanganate, and it is best to use 0.005M *N*-phenylanthranilic acid as indicator (colour change from yellow-green to purple at end point) [J. Mendham, R.C. Denney, J.D. Barnes and M.J.K. Thomas, *Vogel's Quantitative Chemical Analysis*, 6th Edn, Prentice Hall, Harlow, 2000. ISBN 0582226287].

**Ceric sulfate tetrahydrate** is obtained thus: pure ceric ammonium nitrate [16774-21-3] in a small volume of H<sub>2</sub>O is treated with ammonia, Ce(OH)<sub>4</sub> separates, is collected and is dissolved in concentrated H<sub>2</sub>SO<sub>4</sub>. This solution is evaporated to dryness and the residue is crystallised from H<sub>2</sub>O (perhaps a small amount of H<sub>2</sub>SO<sub>4</sub> should be added to avoid hydrolysis). A small volume of H<sub>2</sub>O should be used as boiling with a large volume of H<sub>2</sub>O leads to separation of the basic salt. The yellow-orange powder or orthorhombic crystals lose H<sub>2</sub>O to give the anhydrous salt at 180-200°, but decomposes >350° to give the basic salt CeOSO<sub>4</sub>. [Vanino *Handbuch der Präparativen Chemie* Vol I, 3<sup>rd</sup> Edn, pp 755-756, Stuttgart 1925, Muthmann & Stützel *Chem Ber* 33 1763 1900.]

**Cesium bromide** [7787-69-1] M 212.8, m 636°, b ca 1300°, d<sub>4</sub><sup>20</sup> 4.44. It is very soluble in H<sub>2</sub>O, soluble in EtOH but insoluble in Me<sub>2</sub>CO. Dissolve it in the minimum volume of H<sub>2</sub>O, filter and precipitate it by adding Me<sub>2</sub>CO. Filter off the solid and dry it at 100°. Also recrystallise it from water (0.8ml/g) by partial evaporation in a desiccator.

**Cesium carbonate** [534-17-8] M 325.8, m 792°(at red heat). Crystallise it from ethanol (10ml/g) by partial evaporation. [Dönges in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 988 1963.]

**Cesium chloride** [7647-17-8] M 168.4, m 645°, b 1303°, d<sub>4</sub><sup>20</sup> 3.99. It is soluble in H<sub>2</sub>O but can be purified by crystallisation from H<sub>2</sub>O [solubility in g percent: 162.3(0.7°), 182.2(16.2°) and 290(at bp 119.4°)] and dried in a high vacuum. It is soluble in EtOH and is *deliquescent*; keep it in a tightly closed container. [Dönges in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I pp 951-955 1963.] For further purification of CsCl, a concentrated aqueous solution of the practically pure reagent is treated with an equivalent weight of I<sub>2</sub> and Cl<sub>2</sub> is bubbled into the solution until precipitation of CsCl<sub>2</sub>I is complete. Recrystallisation yields a salt that is free from other alkali metals. It is then decomposed to pure CsCl on heat-

ing. [Harned & Schupp *J Am Chem Soc* **52** 3886 1930.] It can also be recrystallised from acetone/water, or from water (0.5ml/g) by cooling in a CaCl<sub>2</sub>/ice bath. Dry it at 78° under vacuum.

**Cesium chromate** [56320-90-2] **M 381.8, pK<sub>1</sub><sup>25</sup> 0.74, pK<sub>2</sub><sup>25</sup> 6.49 (for H<sub>2</sub>CrO<sub>4</sub>).** Crystallise the chromate from water (1.4ml/g) by partial evaporation in a desiccator. [Boer et al. *Z Anorg Allgem Chem* **191** 113 1930, Hein & Herzog in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **II** p 1389 1963.]

**Cesium fluoride** [13400-13-0] **M 151.9, m 703°.** Crystallise it from aqueous solution by adding ethanol.

**Cesium iodide** [7789-17-5] **M 259.8, m 621°, b~1280°, d<sub>4</sub><sup>20</sup> 4.5.** Crystallise it from warm water (1ml/g) by cooling to -5°.

**Cesium nitrate** [7789-18-6] **M 194.9, m 414°(dec), d<sub>4</sub><sup>20</sup> 3.65.** It crystallises from water (0.6ml/g) between 100° and 0°. After 1 crystallisation of 99.9% grade salt, it had K, Na and Se at 0.8, 0.4 and 0.2 ppm respectively.

**Cesium perchlorate** [13454-84-7] **M 232.4, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>).** Crystallise it from water (4ml/g) between 100° and 0°.

**Cesium sulfate** [10294-54-9] **M 361.9, m 1005°, d<sub>4</sub><sup>20</sup> 4.24.** Crystallise it from water (0.5ml/g) by adding ethanol and cooling.

**Chlorine** [7782-50-5] **M 70.9, m -101.5°, b -34.0°, d<sub>4</sub><sup>20</sup> 2.898.** Pass the gas in succession through aqueous KMnO<sub>4</sub>, dilute H<sub>2</sub>SO<sub>4</sub>, conc H<sub>2</sub>SO<sub>4</sub>, and a drying tower containing Mg(ClO<sub>4</sub>)<sub>2</sub>. Or bubble it through water, dry it over P<sub>2</sub>O<sub>5</sub> and distil it from bulb to bulb in a vacuum line. One volume of water dissolves 4.6 volumes of Cl<sub>2</sub> at 0°, 2.15 volumes at 20°, 1.22 volumes at 50° and 0.39 volumes at 90°. [Schmeisser in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 272 1963.] **HIGHLY TOXIC.**

**Chlorine trifluoride** [7790-91-2] **M 92.5, b 12.1°.** Impurities include chloryl fluoride, chlorine dioxide and hydrogen fluoride. Pass it first through two U-tubes containing NaF to remove HF, then through a series of traps in which the liquid is fractionally distilled. It can be purified *via* the KF complex; KClF<sub>4</sub>, formed by adding excess ClF<sub>3</sub> to solid KF in a stainless steel cylinder in a dry-box and shaking overnight. After pumping out the volatile materials, pure ClF<sub>3</sub> is obtained by heating the bomb to 100-150° and condensing the evolved gas in a -196° trap [Schack et al. *Chem Ind (London)* 545 1967]. It attacks glass very vigorously. **HIGHLY TOXIC.**

**Chlorosulfonic (chlorosulfuric) acid** [7790-94-5] **M 116.5, b 151-152°/750mm, d<sub>4</sub><sup>20</sup> 1.753, n<sub>D</sub><sup>20</sup> 1.4929, pK<sup>25</sup> -5.9 (aqueous H<sub>2</sub>SO<sub>4</sub>).** Distil the acid at atmospheric pressure in an all-glass apparatus, taking the fraction boiling at 156-158°. It reacts **EXPLOSIVELY** with water. **wear gloves and face protection.** [Cremlyn *Chlorosulfonic acid: A Versatile Reagent*, Royal Society of Chemistry UK, 2002, p 308, ISBN 0854044981, Fehér in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 385 1963].

**Chromic chloride (anhydrous)** [10025-73-7] **M 158.4, m 1152°, pK<sub>1</sub><sup>25</sup> 3.95, pK<sub>2</sub><sup>25</sup> 5.55, pK<sub>3</sub><sup>25</sup> 10.5 (for Cr<sup>3+</sup>).** Sublime the chloride in a stream of dry HCl. *Alternatively*, the impure chromic chloride (100g) is added to 1L of 10% aqueous K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> and several millilitres of conc HCl and the mixture is brought to a gentle boil with constant stirring for 10 minutes. (This removes a reducing impurity.) The solid is separated and washed by boiling with successive 1L lots of distilled water until the wash water no longer gives a test for chloride ion, then dry it at 110° [Poulsen & Garner *J Am Chem Soc* **81** 2615 1959, Hein & Herzog in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **II** p 1338 1965].

**Chromium ammonium sulfate dodecahydrate** [34275-72-4 (*hydrate*), 13548-43-1 (*anhydrous*)] **M 478.4, m 94° loses 9H<sub>2</sub>O then dehydrates at 300°, d<sub>4</sub><sup>20</sup> 1.72.** Crystallise the double salt from a saturated aqueous solution at 55° by cooling slowly with rapid mechanical stirring. The resulting fine crystals are filtered on a Büchner funnel, partly dried on a porous plate, then equilibrated for several months in a vacuum desiccator

over crude chromium ammonium sulfate (partially dehydrated by heating at 100° for several hours before use) [Johnson et al. *J Am Chem Soc* **75** 3922 1953].

**Chromium (II) chloride (anhydrous)** [10049-05-5] **M 122.9, m 824°, d<sub>4</sub><sup>14</sup> 2.75**. It is obtained from the *dihydrate* by heating *in vacuo* at 180°. It is a very *hygroscopic* white powder which dissolves in H<sub>2</sub>O to give a sky blue solution. It is stable in dry air but oxidises rapidly in moist air and should be stored in air tight containers. It sublimates at 800° in a current of HCl gas and should be cooled in the presence of HCl gas. *Alternatively*, it can be washed with air-free Et<sub>2</sub>O and dried at 110-120°. [Burg *Inorg Synth III* 150 1950, Balthis & Bailar (4 H<sub>2</sub>O) *Inorg Synth I* 125 1939, Hein & Herzog in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **II** pp 1336-1338 1965.]

**Chromium hexacarbonyl** [13007-92-6] **M 220.1, m 130°(dec), d<sub>4</sub><sup>20</sup> 1.77**. Wash the complex with cold EtOH, then Et<sub>2</sub>O, and allow it to dry in air. *Alternatively*, recrystallise it from dry Et<sub>2</sub>O. This is best accomplished by placing the hexacarbonyl in a Soxhlet extractor and extracting exhaustively with dry Et<sub>2</sub>O. Pure Cr(CO)<sub>6</sub> is filtered off and dried in air. Completely colourless refracting crystals are obtained by sublimation at 40-50°/0.5mm in an apparatus where the collecting finger is cooled by Dry-ice and in which there is a wide short bore between the hot and cold sections to prevent clogging by the crystals. Loss of product in the crystallisation and sublimation is slight. It is important not to overdo the drying as the solid is appreciably volatile and **TOXIC** [vapour pressure is 0.04(8°), 1.0(48°) and 66.5(100°) mm]. Also, do not allow the Et<sub>2</sub>O solutions to stand too long as a brown deposit is formed which is sensitive to light, and to avoid the possibility of violent decomposition. It sinters at 90°, decomposes at 130°, and **EXPLODES** at 210°. [Owen et al. *Inorg Synth III* 156 1950, Podall et al. *J Am Chem Soc* **83** 2057 1961.] **POISONOUS**.

**Chromium (III) nitrate nonahydrate** [7789-02-8] **M 400.2, m ~60° (dec >100°)**. The pure deep violet nonahydrate salt is best prepared freshly from pure recrystallised (3 times) chromium (VI) trioxide and pure nitric acid. A solution is prepared by dissolving 1g of CrO<sub>3</sub> in 3ml of H<sub>2</sub>O and 2ml of pure HNO<sub>3</sub>, and carefully, with stirring (behind a screen) pure MeOH (0.5-1.0ml) is added dropwise carefully (shield) with cooling to avoid a violent reaction. A bluish colour develops as Cr(VI) is reduced to Cr(III). When this solution is diluted to 130,000ppm of the salt, analysis detected by ICP/MS gave the following trace elements (ppm in brackets): Sr (33), Na (3.5), Mo, Ni, Cu, Si and Mg (0.13 each), Se, Zn, Al (0.026 each), Ti (1.5), Fe (4.9), Co (0.07), Sn (0.065), Ba (0.013 and W (0.078). Evaporation of the methanolic solution in high vacuum over CaCl<sub>2</sub> eventually yields pure deep violet rhombic crystals of chromium (III) nitrate *nonahydrate*. An aqueous solution of this salt becomes green on heating but reverts to the violet colour on cooling.

The pale green **anhydrous chromium (III) nitrate**, [13548-38-4] **M 238.0, m dec >60°**, is best obtained pure by mixing a solution of chromium hexacarbonyl in CCl<sub>4</sub> with excess of N<sub>2</sub>O<sub>5</sub> in CCl<sub>4</sub> under N<sub>2</sub> for 12 hours, whereby evolution of gases occurs. Filter off the salt and wash it with CCl<sub>4</sub> in a closed system under N<sub>2</sub>, and dry it *in vacuo*. It is very soluble in H<sub>2</sub>O, EtOAc and Me<sub>2</sub>SO but insoluble in \*C<sub>6</sub>H<sub>6</sub>, CCl<sub>4</sub> and CHCl<sub>3</sub>. The *deliquescent* powder reacts vigorously with Et<sub>2</sub>O. [Addison & Chapman *J Chem Soc* 539 1964.] **Cr(VI)** ions are **CARCINOGENIC** as they cause DNA breaks, and Cr(III) ions affect DNA synthesis.

**Chromium potassium sulfate dodecahydrate** [7788-99-0] **M 499.4, pK<sub>1</sub><sup>25</sup> 0.74, pK<sub>2</sub><sup>25</sup> 6.49 (for H<sub>2</sub>CrO<sub>4</sub>, chromic acid)**. Crystallise it from hot water (2ml/g) by cooling.

**Chromium trioxide (chromic anhydride)** [1333-82-0] **M 100.0, m 197°, dec at 250° to Cr<sub>2</sub>O<sub>3</sub>, d 2.70 (pK<sub>1</sub><sup>25</sup> 0.74, pK<sub>2</sub><sup>25</sup> 6.49, for H<sub>2</sub>CrO<sub>4</sub>, chromic acid)**. It forms red crystals from water (0.5ml/g) between 100° and -5°, or from water/conc HNO<sub>3</sub> (1:5). It separates when potassium or sodium dichromate are dissolved in conc H<sub>2</sub>SO<sub>4</sub>. Dry it in a vacuum desiccator over NaOH pellets. It is a *hygroscopic*, powerful oxidant and can ignite with organic compounds. It is a skin and pulmonary **IRRITANT**. [Keyes et al. *Industrial Chemicals* (Lowenheim & Moran eds.) 4<sup>th</sup> edn J. Wiley pp 270-274 1975.] **CANCER SUSPECT**.

**Chromyl chloride** [14977-61-8] **M 154.9, b 115.7°, d<sub>4</sub><sup>20</sup> 1.911**. Purify it by distillation under reduced pressure. It hydrolyses violently with H<sub>2</sub>O and is a powerful oxidant which explodes with P, and ignites in contact with S, NH<sub>3</sub>, EtOH and many organic compounds. **TOXIC**.

**Claisen alkali (alkali Claisen).** Prepare this from KOH (35g) in H<sub>2</sub>O (25ml) and dilute it to 100ml with MeOH. **STRONGLY CAUSTIC.**

**Cobaltous ammonium sulfate hexahydrate** [13596-46-8] **M 395.5, d<sub>4</sub><sup>20</sup> 1.90.** Crystallise it from boiling water (2ml/g) by cooling. Wash it with ethanol and dry it in a vacuum.

**Cobaltous bromide hexahydrate** [85017-77-2 (xH<sub>2</sub>O), 7789-43-7 (anhydrous)] **M 326.9 (6H<sub>2</sub>O), m 47<sup>o</sup>(dec), b 100<sup>o</sup>(dec), d<sub>4</sub><sup>20</sup> 4.9.** Crystallise it from water (1ml/g) by partial evaporation in a desiccator. The *anhydrous* salt is soluble in EtOH, Me<sub>2</sub>CO, MeOAc to form blue-coloured solutions. [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1517 1965.]

**Cobaltous hexahydrate** [7791-13-1 (6H<sub>2</sub>O), 7646-79-9 (anhydrous)] **M 237.9, m 87<sup>o</sup>(dec), d<sub>4</sub><sup>20</sup> 1.92.** A saturated aqueous solution at room temperature is fractionally crystallised by standing overnight. The first half of the material that crystallises in this way is used in the next crystallisation. The process is repeated several times, water being removed in a dry-box using air filtered through glass wool and dried over CaCl<sub>2</sub> [Hutchinson *J Am Chem Soc* 76 1022 1954]. It has also been crystallised from dilute aqueous HCl. The *hexahydrate* **m 86<sup>o</sup>** forms pink to red deliquescent crystals. It loses 4H<sub>2</sub>O on heating at 52-56<sup>o</sup> and forms the violet *dihydrate* which loses further H<sub>2</sub>O at 100<sup>o</sup> to form the violet *monohydrate* which loses the last H<sub>2</sub>O at 120-140<sup>o</sup> to give the pale blue *anhydrous* deliquescent salt **m 735<sup>o</sup>** and **b 1049<sup>o</sup>**. A pink solution of CoCl<sub>2</sub> in H<sub>2</sub>O becomes blue on heating to 50<sup>o</sup> or adding conc HCl which may precipitate the *mono* or *dihydrate*. The solid *dihydrate* gives a blue-purple solution with EtOH. Note: CoCl<sub>2</sub> in H<sub>2</sub>O is a "sympathetic ink", i.e. writing with an aqueous solution is almost invisible on paper, but becomes blue on warming the paper. On cooling or standing, the writing becomes invisible again. The *anhydrous* salt is soluble in H<sub>2</sub>O, EtOH, Et<sub>2</sub>O, Me<sub>2</sub>CO and pyridine. [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1515 1965.]

**Cobaltous nitrate hexahydrate** [10026-22-9] **M 291.0, m ~55<sup>o</sup>(6H<sub>2</sub>O), 100-105<sup>o</sup>(dec), d<sub>4</sub><sup>20</sup> 1.88.** Crystallise the nitrate from water (1ml/g), or ethanol (1ml/g), by partial evaporation. After 3 crystallisations from H<sub>2</sub>O it contains: metal (ppm) As (8), Fe (1.2), K (1), Mg (4), Mn (4), Mo (4), Na (0.6), Ni (18), Zn (1.6). The *hexahydrate* gives the pink *anhydrous* salt by the action of HNO<sub>3</sub> and N<sub>2</sub>O<sub>5</sub>. The *hexahydrate* melts at ~55<sup>o</sup> to give a red liquid which decomposes on further heating at 100-105<sup>o</sup> to form Co<sub>3</sub>O<sub>4</sub>.

**Cobaltous perchlorate hexahydrate** [13478-33-6] **M 365.9, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>).** Crystallise it from warm water (0.7ml/g) by cooling.

**Cobaltous potassium sulfate** [13596-22-0] **M 329.4.** Crystallise it from water (1ml/g) between 50<sup>o</sup> and 0<sup>o</sup>, and dry it in a vacuum desiccator over conc H<sub>2</sub>SO<sub>4</sub>.

**Cobaltous sulfate heptahydrate** [10026-24-1 (7H<sub>2</sub>O), 60459-08-7 (xH<sub>2</sub>O), 10124-43-3 (anhydrous)] **M 281.1, m (see text), d<sub>4</sub><sup>20</sup> 2.03.** Crystallise it three times from conductivity water (1.3ml/g) between 100<sup>o</sup> and 0<sup>o</sup> depending on which hydrate is required. The *heptahydrate* crystallises below 44<sup>o</sup> and is efflorescent with **m 97<sup>o</sup>**. Between 44<sup>o</sup> and 70<sup>o</sup> the monoclinic *hexahydrate* CoSO<sub>4</sub>·6H<sub>2</sub>O **m 41.5<sup>o</sup>** is formed, and above 70<sup>o</sup> the *monohydrate* CoSO<sub>4</sub>·H<sub>2</sub>O **m 71<sup>o</sup>** is obtained. The pale reddish or lavender-coloured *anhydrous* salt is obtained by heating the hydrate above 250<sup>o</sup>, boiling with conc H<sub>2</sub>SO<sub>4</sub> or by heating with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>.

**Cupric ammonium chloride dihydrate** [10534-87-9 (hydrate), 15610-76-1 (anhydrous)] **M 277.5, m 110-120<sup>o</sup>(anhydrous) then dec at higher temperature, d<sub>4</sub><sup>20</sup> 2.0.** Crystallise it from weak aqueous HCl (1ml/g).<sup>\*</sup> It crystallises out of a hot solution of CuCl<sub>2</sub> saturated with NH<sub>3</sub> gas.

**Cupric bromide** [7789-45-9] **M 223.4, m 498<sup>o</sup>, b 900<sup>o</sup>, d<sub>4</sub><sup>20</sup> 4.7.** Crystallise it twice by dissolving it in water (140ml/g), filtering to remove any Cu<sub>2</sub>Br<sub>2</sub>, and concentrating under vacuum at 30<sup>o</sup> until crystals appear. The cupric bromide is then allowed to crystallise by leaving the solution in a vacuum desiccator containing P<sub>2</sub>O<sub>5</sub> [Hope et al. *J Chem Soc* 5226 1960, Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1009 1965].

**Cupric chloride** [7447-39-4] **M 134.4, m 498°, 630°(dec)**. Crystallise the chloride from hot dilute aqueous HCl (0.6ml/g) by cooling in a CaCl<sub>2</sub>-ice bath. It is dehydrated by heating on a steambath under vacuum. It is deliquescent in moist air but efflorescent in dry air. The *dihydrate* is emerald green but blue when free from solvent. Concentrated solutions are yellow-green in colour but are blue when free from solvent. Concentrated solutions are yellow-green and become yellow on adding conc HCl. A very dilute solution is pure blue due to Cu(H<sub>2</sub>O)<sub>4</sub><sup>2+</sup> [Donan & Bassett *J Chem Soc* **81** 939 1902]. CuCl<sub>2</sub> is very deliquescent and is soluble in MeOH or EtOH to give green crystals of Cu(ROH)<sub>2</sub>Cl<sub>2</sub>. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1008 1965.]

**Cupric nitrate trihydrate** [10031-43-3 (3H<sub>2</sub>O), 3251-23-8 (anhydrous)] **M 241.6, m 114°, b 170°(dec), d<sub>4</sub><sup>20</sup> 2.0**. Crystallise it from weak aqueous HNO<sub>3</sub> (0.5ml/g) by cooling from room temperature. The *anhydrous* salt can be prepared by dissolving copper metal in a 1:1 mixture of liquid NO<sub>2</sub> and ethyl acetate and purified by sublimation [Evans et al. *J Chem Soc, Faraday Trans 1* **75** 1023 1979]. The *hexahydrate* dehydrates to the *trihydrate* at 26°, and the *anhydrous salt* sublimes between 150 and 225°, but melts at 255-256° and is *deliquescent*.

**Cupric perchlorate hexahydrate** [10294-46-9 (hydrate), 13770-18-8] **M 370.5, m 230-240°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Crystallise it from distilled water. The *anhydrous salt* is hygroscopic.

**Cupric sulfate (blue vitriol, bluestone)** [7758-98-7] **M 159.6, m >560°**. After adding 0.02g of KOH to a litre of nearly saturated aqueous solution of the sulfate, it is left for two weeks, then the precipitate is filtered on to a fibreglass filter with pore diameter of 5-15 microns. The filtrate is heated to 90° and allowed to evaporate until some CuSO<sub>4</sub>.5H<sub>2</sub>O crystallises out. The solution is then filtered hot and cooled rapidly to give crystals which are freed from mother liquor by filtering under suction [Geballe & Giauque *J Am Chem Soc* **74** 3513 1952]. *Alternatively*, crystallise the sulfate from water (0.6ml/g) between 100 and 0°. The *pentahydrate* is slowly efflorescent, losing 2H<sub>2</sub>O at 30°, two more H<sub>2</sub>O are lost at 110° and a white anhydrous powder (desiccant) is obtained on heating above 250°.

**Cuprous bromide** [7787-70-4] **M 143.4, m 497°, b 1345°, d<sub>4</sub><sup>20</sup> 4.72**. Purify it as for cuprous iodide but using aqueous NaBr. [Keller & Wycoff *Inorg Synth* **II** 3 1946, Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1006 1965.]

**Cuprous chloride** [7758-89-6] **M 99.0, m 430°, b~1400°**. Dissolve it in strong HCl, precipitate it by diluting with water and filter it off. Wash the solid with ethanol and diethyl ether, then dry it and store it in a vacuum desiccator [Österlöf *Acta Chem Scand* **4** 375 1950]. *Alternatively*, to an aqueous solution of CuCl<sub>2</sub>.2H<sub>2</sub>O is added, with stirring, an aqueous solution of anhydrous sodium sulfite. The colourless product is dried at 80° for 30 minutes and stored under N<sub>2</sub>. Cu<sub>2</sub>Cl<sub>2</sub> can be purified by zone-refining [Hall et al. *J Chem Soc, Faraday Trans 1* **79** 343 1983]. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1005 1965.]

**Cuprous cyanide** [544-92-3] **M 89.6, m 474°**. Wash the cyanide thoroughly with boiling H<sub>2</sub>O, then with EtOH. Dry it at 100° to a fine soft powder. It dissolves in excess alkali cyanide solutions to form the very soluble complex ion Cu(CN)<sub>4</sub><sup>3-</sup>. [Bassett & Corbett *J Chem Soc* **125** 1660 1924, Barber *J Chem Soc* 79 1943.]

**Cuprous iodide** [7681-65-4] **M 190.5, m 605°, b 1336°, d<sub>4</sub><sup>25</sup> 5.63**. It can be freshly prepared by dissolving an appropriate quantity of CuI in boiling saturated aqueous NaI over 30 minutes. Pure CuI is obtained by cooling and diluting the solution with water, followed by filtering and washing sequentially with H<sub>2</sub>O, EtOH, EtOAc, Et<sub>2</sub>O and pentane, then drying *in vacuo* for 24 hours [Dieter, *J Am Chem Soc* **107** 4679 1985]. *Alternatively*, wash it with H<sub>2</sub>O, then EtOH and finally with Et<sub>2</sub>O containing a little iodine. Traces of H<sub>2</sub>O are best removed first by heating at 110° and then at 400°. Excess of I<sub>2</sub> is removed completely at 400°. It dissolves in Et<sub>2</sub>O if an amine is present to form the amine complex. On heating it becomes red, then black, but changes to white on cooling. It is sparingly soluble in H<sub>2</sub>O or alkali iodide solutions but readily soluble in NH<sub>3</sub> (which absorbs CO) and in cyanide or thiosulfate solutions. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1007 1965, Bawn & Ledwith *Chem Ind (London)* 1180 1957.]

**Cuprous thiocyanate** [18223-42-2] **M 121.6, pK<sup>25</sup> -1.85 (for HSCN).** Purify it as for cuprous iodide but using aqueous NaSCN. [Demmerle et al. *Ind Eng Chem* **42** 2 1950, Newman *Analyst* **88** 500 1963.]

**Cyanamide** [420-04-2] **M 42.0, m 43°, 45°, 46°, b 85-87°/0.5mm, pK<sub>1</sub><sup>20</sup> -0.36 (1.1 at 29°), pK<sub>2</sub><sup>20</sup> 10.27.** Purify it by placing *ca* 15g in a Soxhlet thimble and extracting exhaustively (2-3 hours) with two successive portions of Et<sub>2</sub>O (400ml, saturated with H<sub>2</sub>O by shaking before use) containing two drops of 1N acetic acid. Two successive portions of Et<sub>2</sub>O are used so that the NH<sub>2</sub>CN is not heated for too long. Each extract is dried over Na<sub>2</sub>SO<sub>4</sub> (30g), then combined and evaporated under reduced pressure. The NH<sub>2</sub>CN may be stored unchanged at 0° in Et<sub>2</sub>O solution in the presence of a trace of AcOH. Extracts from several runs may be combined and evaporated together. The residue from evaporation of an Et<sub>2</sub>O solution is a colourless viscous oil which sets to a solid and can be recrystallised from a mixture of 2 parts of \*C<sub>6</sub>H<sub>6</sub> and 1 part of Et<sub>2</sub>O. Concentrating an aqueous solution of NH<sub>2</sub>CN at high temperatures causes **EXPLOSIVE** polymerisation. [Kurzer & Lawson *Org Synth Coll Vol IV* 645 1963, Pinck & Salisbury *Inorg Synth III* 39 1950, Soloway & Lipschitz *J Org Chem* **23** 613 1958.] *Hygroscopic.* [Beilstein **3** IV 145.]

**Cyanogen bromide** [506-68-3] **M 105.9, m 49-51°, b 60-62°/atm.** *All operations with this substance should be performed in a very efficient fume cupboard—it is very POISONOUS and should be handled in small amounts. Fresh commercial material is satisfactory for nearly all purposes and does not need to be purified. It is a white crystalline solid with a strong cyanide odour. If it is reddish in colour and partly liquid or paste-like, then it is too far gone to be purified, and fresh material should be sought.* It can be purified by distillation using small amounts at a time, and using a short wide-bore condenser because it readily solidifies to a crystalline white solid which may clog the condenser. *An appropriate gas mask should be used when transferring the molten solid from one container to another, and the operation should be done in an efficient fume cupboard.* The melting point (**m 49-51°**) should be measured in a sealed tube. [Hartman & Dreger *Org Synth Coll Vol II* 150 1948.]

**Cyanogen iodide** [506-78-5] **M 152.9, m 146-147°.** *This compound is POISONOUS, and the precautions for cyanogen bromide (above) apply here.* The reagent (*ca* 5.9g) is dissolved in boiling CHCl<sub>3</sub> (15ml), filtered through a plug of glass wool into a 25ml Erlenmeyer flask. Cool to room temperature for 15 minutes, then place it in an ice-salt bath and cool to -10°. This cooling causes a small aqueous layer to separate as ice. The ice is filtered with the CNI, but melts on the filter and is also removed with the CHCl<sub>3</sub> used as washing liquid. The CNI which is collected on a sintered glass funnel is washed 3x with CHCl<sub>3</sub> (1.5ml at 0°) and freed from last traces of solvent by placing it on a watch glass and exposing it to the atmosphere in a good fume cupboard at room temperature for 1 hour to give colourless needles (*ca* 4.5g), **m 146-147°** (sealed capillary totally immersed in the oil bath). The yield depends slightly on the rapidity of the operation; in this way loss by sublimation can be minimised. If desired, it can be sublimed under reduced pressure at temperatures at which CNI is only slowly decomposed into I<sub>2</sub> and (CN)<sub>2</sub>. The vacuum will need to be renewed constantly due to the volatility of CNI. [Bak & Hillebert *Org Synth Coll Vol IV* 207 1963.]

**Decaborane** [17702-41-9] **M 122.2, m 99.7-100°, b 100°/19mm, 213°/atm.** Purify decaborane by vacuum sublimation at 80°/0.1mm, followed by crystallisation from methylcyclohexane, CH<sub>2</sub>Cl<sub>2</sub>, or dry olefin-free-*n*-pentane, the solvent being subsequently removed by storing the crystals in a vacuum desiccator containing CaCl<sub>2</sub>. It is soluble in H<sub>2</sub>O but is slowly decomposed to give H<sub>2</sub>. It is soluble in alkali, and on acidification it liberates H<sub>2</sub>. **TOXIC.** [Greenwood in *Comprehensive Chemistry (Ed Bailar et al.)* Pergamon Press Vol **1** pp 818-837 1973.]

**Deuterium** [7782-39-0] **M 4.** Pass the gas over activated charcoal at -195° [MacIver & Tobin *J Phys Chem* **64** 451 1960]. Purify it also by diffusion through nickel [Pratt & Rogers, *J Chem Soc, Faraday Trans I* **92** 1589 1976]. Always check deuterium for radioactivity to determine the amount of tritium in it (see D<sub>2</sub>O below).

**Deuterium oxide** [7789-20-0] **M 20, f 3.8°/760mm, b 101.4°/760mm, d<sub>4</sub><sup>20</sup> 1.105.** Distil it from alkaline KMnO<sub>4</sub> [de Giovanni & Zamenhof *Biochem J* **92** 79 1963]. **NOTE that D<sub>2</sub>O invariably contains tritiated water and will therefore be RADIOACTIVE; always check the radioactivity level of D<sub>2</sub>O in a scintillation**

counter before using.

**cis-Diamminedichloroplatinum(II) (Cisplatin)** [15663-27-1] **M 300.1, m 270°(dec)**. Recrystallise it from dimethylformamide and check the purity by IR and UV-VIS spectroscopy. [Raudaschl et al. *Inorg Chim Acta* **78** 143 1983.] **HIGHLY TOXIC, SUSPECTED CARCINOGEN**.

**Diammonium hydrogen orthophosphate** [7783-28-0] **M 132.1**. Crystallise it from water (1ml/g) between 70 and 0°. Its solubility in H<sub>2</sub>O is 59% at room temperature and 200% at boiling point. It slowly evolves NH<sub>3</sub> and should be stored in a well-stoppered container. After one crystallisation, ACS grade salt had Fe, Mo, Na, Se and Ti at 1, 0.2, 1.4, 0.2 and 0.8ppm, respectively. [Gmelin's, Ammonium (8th edn) **23** pp422-426 1936.]

**Dinitrogen tetroxide (nitrogen dioxide, N<sub>2</sub>O<sub>4</sub>)** [10544-72-6] **M 92.0 m -11.2°, b 21.1°**. Purify it by oxidation at 0° in a stream of oxygen until the blue colour changes to red-brown. *Alternatively* distil it from P<sub>2</sub>O<sub>5</sub>, then solidify it by cooling in a deep-freeze (at -78°, giving nearly colourless crystals). Oxygen can be removed by alternate freezing and melting. **TOXIC VAPOUR**. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 488-489 1963.]

**Disodium hydrogen orthophosphate (anhydrous)** [7558-79-4] **M 142.0 (see pK of H<sub>3</sub>PO<sub>4</sub>)**. Crystallise the salt twice from warm water, by cooling. Dry in air, then in an oven overnight at 130°. It should be dried before use as it is slightly *hygroscopic*. It forms di-, hepta- and deca- hydrates.

**Ferric Bromide** [10031-26-2] **M 395.6, m >130°(dec)**. Sublime it in a sealed tube with Br<sub>2</sub> at 120°-200°. [Lux in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1494 1965.]

**Ferric chloride (anhydrous)** [7705-08-0] **M 162.2, m >300°(dec)**. Sublime it at 200° in an atmosphere of chlorine. It is an "iron-black" coloured powder with green iridescence. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, EtOH, Et<sub>2</sub>O, Me<sub>2</sub>CO and pyridine in which the monomeric species "FeCl<sub>3</sub>" exist whereas in the vapour state it exists in the dimeric "Fe<sub>2</sub>Cl<sub>6</sub>" form. Store it in a weighing bottle inside a desiccator as it absorbs moisture from air to form the yellow *hexahydrate* (see next entry). [Tarr *Inorg Synth* **III** 191 1950, Pray *Inorg Synth* **V** 153 1957, Epperson *Inorg Synth* **VII** 163 1963.]

**Ferric chloride hexahydrate** [10025-77-1] **M 270.3, m 37°(dec), pK<sub>1</sub><sup>25</sup> 2.83, pK<sub>2</sub><sup>25</sup> 4.59 (for hydrolysis of Fe<sup>3+</sup>)**. An aqueous solution, saturated with the salt at room temperature, is cooled to -20° for several hours. Separation of the crystals is slow, even with scratching and seeding, and it is generally necessary to stir this overnight. The presence of free HCl retards crystallisation. [Linke *J Phys Chem* **60** 91 1956].

**Ferric nitrate nonahydrate** [7782-61-8] **M 404.0, m 47°(dec)**. It crystallises from aqueous solutions of moderately strong HNO<sub>3</sub> as the pale violet *nonahydrate* **m 40°** and is soluble in EtOH and Me<sub>2</sub>CO. With more concentrated aqueous solutions (containing some HNO<sub>3</sub>), the *hexahydrate* crystallises out **m 60.5°**. The *anhydrous* salt is slightly deliquescent and decomposes at 47°. [Lambert & Thomson *J Chem Soc* **97** 2426 1920, Gmelin's, *Iron* (8th edn) **59** Part B pp 161-172 1932.]

**Ferric perchlorate nonahydrate** [13537-24-1] **M 516.3, pK<sub>1</sub><sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Crystallise it twice from conc HClO<sub>4</sub>, the first time in the presence of a small amount of H<sub>2</sub>O<sub>2</sub> to ensure that the iron is fully oxidised [Sullivan *J Am Chem Soc* **84** 4256 1962]. Extreme care should be taken with this preparation because it is potentially **EXPLOSIVE**.

**Ferric sulfate x-hydrate** [10028-22-5] **M 399.9 + xH<sub>2</sub>O**. Dissolve the sulfate in the minimum volume of dilute aqueous H<sub>2</sub>SO<sub>4</sub> and allow it to evaporate at room temperature until yellowish-white crystals start to form. Do not concentrate by boiling off the H<sub>2</sub>O as basic salts will be formed. Various *hydrates* are formed; the common ones are the dodeca and nona *hydrates* which are violet in colour. The *anhydrous salt* is colourless

and is quite *hygroscopic*, but it dissolves in H<sub>2</sub>O slowly unless ferrous sulfate is added. [*Gmelin's, Iron* (8th edn) pp 439-462 1932.]

**Ferrous bromide** [20049-65-4] **M 215.7 + xH<sub>2</sub>O, m 684°, d<sup>25</sup> 4.63.** It crystallises from air-free H<sub>2</sub>O to provide the *hexahydrate* as pale green to bluish-green rhombic prisms. On heating at 49° H<sub>2</sub>O is lost and the *tetrahydrate* is formed. On further heating at 83° more H<sub>2</sub>O is lost and the *dihydrate* is formed as a light yellow to dark brown *hygroscopic* powder. The ferrous iron in aqueous solutions of these salts readily oxidise to ferric iron. The salts should be stored over H<sub>2</sub>SO<sub>4</sub> under N<sub>2</sub> in tightly closed containers. They have some solubility in EtOH. [Baxter *Z Anorg Chem* **38** 236 1904, Kühn & Ernst *Z Anorg Allgem Chem* **317** 84 1962, Lux in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1493 1965.]

**Ferrous chloride tetrahydrate** [13478-10-9] **M 198.8, m 105°(dec), pK<sub>1</sub><sup>25</sup> 6.7, pK<sub>2</sub><sup>25</sup> 9.3 (for aquo Fe<sup>2+</sup>).** A 550ml round-bottomed Pyrex flask is connected, *via* a glass tube fitted with a medium porosity sintered-glass disc, to a similar flask. To 240g of FeCl<sub>2</sub>·4H<sub>2</sub>O in the first flask is added conductivity water (200ml), 38% HCl (10ml), and pure electrolytic iron (8-10g). A stream of purified N<sub>2</sub> gas is passed through the assembly, escaping through a mercury trap. The salt is dissolved by heating which is continued until complete reduction has occurred. By inverting the apparatus and filtering (under N<sub>2</sub> pressure) through the sintered glass disc, unreacted iron is removed. After cooling and crystallisation, the unit is again inverted, and the crystals of ferrous chloride are filtered free from mother liquor by applied N<sub>2</sub> pressure. Partial drying by overnight evacuation at room temperature gives a mixed hydrate which, on further evacuation on a water bath at 80°, loses water of hydration and absorbed HCl (with vigorous effervescence) to give a white powder of FeCl<sub>2</sub>·2H<sub>2</sub>O (see below). [Gayer & Wootner *J Am Chem Soc* **78** 3944 1956, (2H<sub>2</sub>O) Gayer & Wootner *Inorg Synth* **V** 179 1957.]

**Ferrous chloride** [7758-94-3] **M 126.8, m 674°, b 1023°, d<sup>25</sup> 3.16.** It sublimes in a stream of HCl at *ca* 700°, or in H<sub>2</sub> below 300°. Its vapour pressure at 700° is 12mm. It forms white *hygroscopic* rhombohedral crystals with a green tint which oxidise in air to FeCl<sub>3</sub> and Fe<sub>2</sub>O<sub>3</sub>. It is soluble in H<sub>2</sub>O, EtOH Me<sub>2</sub>CO but insoluble in Et<sub>2</sub>O. The *tetrahydrate* is pale green to pale blue in colour and loses 2H<sub>2</sub>O at 105-115°. The *dihydrate* loses H<sub>2</sub>O at 120°. [Anhydrous FeBr<sub>2</sub> can be obtained by carefully dehydrating the *tetrahydrate* in a stream of HBr and N<sub>2</sub>, and it can be sublimed under N<sub>2</sub>.] The ferrous iron in aqueous solutions of these salts readily oxidises to ferric iron. (See above.) [Kovacuumic & Brace *Inorg Synth* **VI** 172 1960, Lux in *Handbook of Preparative Inorganic Chemistry* (Ed Brauer) Academic Press Vol **II** p 1491 1965.]

**Ferrous perchlorate hexahydrate** [13933-23-8] **M 362.9, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>).** Crystallise it from HClO<sub>4</sub>. [CARE, see ferric perchlorate above.]

**Ferrous sulfate heptahydrate (green vitriol)** [7782-63-0] **M 278.0, m ~60°(dec).** Crystallise the sulfate from 0.4M H<sub>2</sub>SO<sub>4</sub>, or precipitate it from an aqueous solution with EtOH. It is *efflorescent* in dry air, and is converted to the *tetrahydrate* at 57°, then to the *monohydrate* at 65° (or by heating the heptahydrate in a vacuum at 140°). It forms a brown-black complex, FeSO<sub>4</sub>·NO, with nitric oxide and is used in a qualitative test for nitrates ("brown ring" test).

**Fluorine** [7782-41-4] **M 38.0, b -129.2°.** Pass the gas through a bed of NaF at 100° to remove HF and SiF<sub>4</sub>. [For description of stills used in fractional distillation, see Greenberg et al. *J Phys Chem* **65** 1168 1961; Stein et al. *Purification of Fluorine by Distillation, Argonne National Laboratory, ANL-6364* 1961 (from Office of Technical Services, US Dept of Commerce, Washington 25).] **HIGHLY TOXIC.**

**Fluoroboric acid (tetrafluoroboric acid, fluoboric acid)** [16872-11-0] **M 87.8, b 130°(dec), pK<sup>25</sup> -4.9.** Crystallise fluoroboric acid several times from conductivity water. It can be stored in a glass vessel at room temperature. It is available commercially as ~48% aqueous solution. It is most useful for preparing *tetrafluoroborate salts* which are generally insoluble. For example, addition of the acid to aryldiazonium salt solutions precipitates the more stable aryldiazonium tetrafluoroborate salts which can be washed with H<sub>2</sub>O to remove impurities, followed by EtOH and Et<sub>2</sub>O, and stored for short periods of time before further use. It is a catalyst for preparing acetals. [Mathers et al. *J Am Chem Soc* **37** 1516 1915, Wamser & Christian *J Am Chem*

*Soc* 70 1209 1940, Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 221-222 1963.] **STRONG IRRITANT and is TOXIC.**

**Gallium** [7440-55-3] **M 69.7, m 29.78°, b 2403°, d<sup>29.6</sup> 5.904, d<sup>29.8</sup> 6.095.** Dissolve the metal in dilute HCl and extract it with Et<sub>2</sub>O. Bubbling H<sub>2</sub>S through the solution removes many metals, and a second extraction with Et<sub>2</sub>O frees Ga further from metal impurities, except for Mo, Th(III) and Fe which are largely removed by precipitation with NaOH. The solution is then electrolysed in 10% NaOH with a Pt anode and cathode (2-5A at 4-5V) to deposit Ga, In, Zn and Pb, from which Ga was obtained by fractional crystallisation of the melt [Hoffman *J Res Nat Bur Stand* 13 665 1934]. Ga is also purified by heating to boiling in 0.5-1M HCl, then heating to 40° in water and pouring the molten Ga with water under vacuum through a glass filter (30-50 μ pore size), to remove any unmelted metals or oxide film. The Ga is then fractionally crystallised from the melt under water. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 837 1963.]

**Gallium (III) Chloride** [13450-90-3] **M 176.1, m 77.8°, b 133°/100mm, 197.7°/700mm, d<sub>4</sub><sup>20</sup> 2.47, pK<sub>1</sub><sup>25</sup> 2.91, pK<sub>2</sub><sup>25</sup> 3.70, pK<sub>3</sub><sup>20</sup> 4.42 (for Ga<sup>3+</sup>).** The pure compound can be obtained by redistillation in a stream of Cl<sub>2</sub> or Cl<sub>2</sub>/N<sub>2</sub> followed by vacuum sublimation or zone refining. It forms colourless needles which give *gallium dichloride* [Ga(GaCl<sub>4</sub>), m 172.4°] on heating. It dissolves in H<sub>2</sub>O with liberation of heat. It is soluble in Et<sub>2</sub>O and can be extracted from an HCl solution with Et<sub>2</sub>O. [Laubengayer & Schirmer *J Am Chem Soc* 62 1579 1940, Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 846 1963.]

**Gallium (III) nitrate nonahydrate** [63462-65-7] **M 417.9, m ca 65°.** Recrystallise the nitrate from H<sub>2</sub>O (solubility is 295g/100ml at 20°). It forms a white *deliquescent*, colourless powder soluble in H<sub>2</sub>O, absolute EtOH and Et<sub>2</sub>O. It loses HNO<sub>3</sub> upon heating at 40°. Addition of Et<sub>2</sub>O to a warm ethanolic solution (40-50°) of Ga(NO<sub>3</sub>)<sub>3</sub>·9H<sub>2</sub>O precipitates Ga(OH)<sub>2</sub>NO<sub>3</sub>·Ga(OH)<sub>3</sub>·2H<sub>2</sub>O. If the salt has partly hydrolysed, dissolve it in conc HNO<sub>3</sub>, reflux, dilute with H<sub>2</sub>O and concentrate on a sand bath. Wash the solid several times by adding H<sub>2</sub>O and evaporating until there is no odour of acid. Dilute the residue to a Ga concentration of 26g/100ml. At this concentration, *spongy* Ga(NO<sub>3</sub>)<sub>3</sub>·xH<sub>2</sub>O separates from the viscous solution. After standing for several days the crystals are collected and dried in a stream of dry air first at room temperature, then at 40°. *Dehydration* is complete after 2 days. Recrystallise it from H<sub>2</sub>O and dry it at water pump vacuum at room temperature. [Reimmann & Tanner *Z Naturforsch* 20B 71 1965, Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 856 1963.]

**Gallium (III) sulfate** [13494-91-2 (*anhydrous*), 13780-42-2 (*hydrate*)] **M 427.6.** Recrystallisation from H<sub>2</sub>O gives the 16-18H<sub>2</sub>O hydrate (solubility at 20° is 170g/100ml). *Alternatively*, dissolve it in 50% H<sub>2</sub>SO<sub>4</sub> and evaporate (60-70°), cool and precipitate it by adding EtOH/Et<sub>2</sub>O. On heating at 165° it provides the *anhydrous* salt, which is a white *hygroscopic* solid. [Reimmann & Tanner *Z Naturforsch* 20B 71 1965.]

**Germanium** [7440-56-4] **M 72.6, m 937°, 925-975°, b 2700°, d<sub>4</sub><sup>20</sup> 5.3.** Copper contamination on the surface and in the bulk of single crystals of Ge can be removed by immersion in molten alkali cyanide under N<sub>2</sub>. The Ge is placed in dry K and/or Na cyanide powder in a graphite holder in a quartz or porcelain boat. The boat is then inserted into a heated furnace which, after a suitable time, is left to cool to room temperature. At 750°, a 1mm thickness of metal requires about 1 minute, whereas 0.5cm needs about half hour. The boat is removed from the furnace, and the solid samples are taken out with plastic-coated tweezers, carefully rinsed in hot water and dried in air [Wang *J Phys Chem* 60 45 1956, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 712 1963]. **Care with the use of cyanide.**

**Germanium (IV) oxide** [1310-53-8] **M 104.6, m 1080°(soluble form), d<sup>25</sup> 6.239; m 1116°(insoluble form) d<sup>25</sup> 4.228, pK<sub>1</sub><sup>25</sup> 9.02, pK<sub>2</sub><sup>25</sup> 12.82 (for germanic acid H<sub>2</sub>GeO<sub>3</sub>).** The oxide (GeO<sub>2</sub>) is usually prepared by hydrolysing redistilled GeCl<sub>4</sub> and igniting it in order to remove H<sub>2</sub>O and chloride. It can be further purified by dissolving in hot H<sub>2</sub>O (solubility is 4g/L cold) evaporating and drying the residual crystalline solid. When the *soluble* form (which is produced in H<sub>2</sub>O at 355°) is heated for 100 hours, it is converted to the *insoluble* form.

This form is stable at temperatures up to 1033°, and fusion at 1080° for 4 hours causes complete de-vitrification and it reverts to the *soluble* form. [Müller & Blank *J Am Chem Soc* **46** 2358 1924, Dennis & Laubengayer *J Am Chem Soc* **47** 1945 1925, Laubengayer & Morton *J Am Chem Soc* **54** 2303 1932, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 706 1963.]

**Germanium tetrabromide** [13450-92-5] M 392.2, m 26°, b 185.9°/atm,  $d_4^{29}$  3.123. Purify it by simple distillation or fractionation depending on purity. It is soluble in EtOH, CHCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub> and Et<sub>2</sub>O. It fumes in moist air and is readily hydrolysed by water. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 718 1963]. **LACHRYMATORY.**

**Germanium tetrachloride** [10038-98-9] M 214.4, m -49.5° (α), -52.0° (β), b 83.1°/760mm, 86.5°/760mm corr,  $d_4^{20}$  1.84. Traces of Cl<sub>2</sub> and HCl can be removed from the liquid by blowing dry air through it for a few hours at room temperature or by shaking it with Hg or Hg<sub>2</sub>Cl<sub>2</sub> and then fractionating it in a vacuum. It decomposes on heating at 950°. It has a sharp penetrating odour and fumes in moist air to give a chalky coat of GeO<sub>2</sub>. It is slowly hydrolysed by H<sub>2</sub>O to give GeO<sub>2</sub>, but distils from conc HCl. [Foster et al. *Inorg Synth* **II** 109 1946, Dennis & Hance *J Am Chem Soc* **44** 304 1922, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 715 1963.] **LACHRYMATORY.**

**Glass powder** (100-300 mesh). Washed with 10% HNO<sub>3</sub>, water and dry in air.

**Gold powder** [7440-57-5] M 196.97, m 1064.79°, b 2808° (2700°),  $d_4^{20}$  19.3, resistivity at 0° is 2.05 μΩ-cm. Massive gold reacts very slowly and it is best converted to the powder which is more readily reactive. While working in a well vented fume cupboard, solid gold (10g, 51mmol) is dissolved in *Aqua Regia*, made from concentrated hydrochloric acid (12M HCl, 50ml) and concentrated nitric acid (16M HNO<sub>3</sub>, 13ml), by heating to 90° which will hasten dissolution, but will require topping up with *Aqua Regia*. Alternatively, the suspension is allowed to stand at ~25° until the solid has dissolved. The solution is evaporated to ca 10ml, concentrated hydrochloric acid (30ml) is added and evaporation is repeated. The solution is diluted with H<sub>2</sub>O (20ml), filtered through a medium sintered-glass funnel, and the filter is washed with small volumes of H<sub>2</sub>O. The filtrate and washings, which contain HAuCl<sub>4</sub>, are heated to boiling and hydroquinone (10g, 91mmol) in hot H<sub>2</sub>O (100ml) is slowly added. The mixture is kept at 90° for 1 hour (sufficient time to reduce all the HAuCl<sub>4</sub> to Au), cooled and filtered through an extraction thimble. The thimble is placed in a Soxhlet apparatus and extracted with MeOH. After 15 minutes, when the circulating MeOH in the thimble is colourless, the thimble is removed and the contents are collected and dried in air to give Au powder quantitatively. [Block *Inorg Synth* **4** 15 1953.]

**Gold (III) bromide (gold tribromide)** [10294-28-7] M 436.7, m 150°(dec). Purify it by adding pure Br<sub>2</sub> to the dark powder, securely stopper the container, warm a little and shake while keeping away from light for ca 48 hours. Remove the stopper and place it over NaOH until free Br<sub>2</sub> is no longer in the apparatus (48-60 hours). The bright yellow needles of the tribromide are stable over NaOH in the dark. It is soluble in H<sub>2</sub>O and in EtOH where it is slowly reduced. Keep it in a cooled, closed container and protect it from light as decomposition causes free gold to be formed. *Auribromic acid* can be obtained by adding the calculated amount of conc HBr to AuBr<sub>3</sub> (actually Au<sub>2</sub>Br<sub>6</sub>) until all dissolves, whereby the acid crystallises out as HAuBr<sub>4</sub>·5H<sub>2</sub>O; a *deliquescent* solid soluble in EtOH with m ca 27°, and store it as above. [Gibson & Colles *J Chem Soc* 2411 1931, Burawoy & Gibson *J Chem Soc* 217 1935, Burawoy & Gibson *J Chem Soc* 219 1935.]

**Gold (I) chloride [aurous chloride, Au(I)Cl]** [10294-29-8] M 232.4, m 289°(dec),  $d^{25}$  7.4. It is best prepared by heating AuCl<sub>3</sub>·3H<sub>2</sub>O, or HAuCl<sub>4</sub> (see next entry) at 100° in a high vacuum until the vapour pressure drops considerably when most of the H<sub>2</sub>O has been expelled. As the HAuCl<sub>4</sub> liquefies again in this process much spluttering will occur. The solid residue (mostly anhydrous AuCl<sub>3</sub>) is then heated to 156° (bromobenzene bath) and further to 170-205° when decomposition to AuCl is complete. Another preparation includes heating AuCl<sub>3</sub> in air at 185° [Thomsen *J Prakt Chem* **13** 337 1876], but purer salt is obtained by heating AuCl<sub>3</sub> in a stream of dry HCl at 175° [Diemer *J Am Chem Soc* **35** 552 1913]. It forms pale yellow crystals which are not deliquescent, slightly soluble in cold EtOH, dissolve in alkali chloride solutions to form chloroaurates(I), e.g. KAuCl<sub>2</sub>, but it decompose in H<sub>2</sub>O to give gold and hydrolysed Au(III) species. [Blitz & Wein *Z Anor Allgem Chem* **148** 192 1925, Glemser & Sauer *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic

Press Vol II p 1055 1965.]

**Gold (III) chloride (hydrate)** [16903-35-8] **M 339.8 + xH<sub>2</sub>O, x ~3, m 229°, b 354°(dec), d<sub>4</sub><sup>20</sup> 3.9.** It is obtained as a dark red crystalline mass by dissolving Au in *aqua regia* and evaporating. When sublimed at 180°, the anhydrous crystals are ruby red. The anhydrous salt is *hygroscopic*, soluble in H<sub>2</sub>O but sparingly soluble in EtOH and Et<sub>2</sub>O. *Aurichloric acid* (chloroauric acid, HAuCl<sub>4</sub>) is formed when AuCl<sub>3</sub> is dissolved in HCl. [Diemer *J Am Chem Soc* **35** 553 1913, Block *Inorg Chem* **4** 14 1953, Glemser & Sauer *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1056 1965.]

**Gold (I) cyanide** [506-65-0] **M 223.0, m dec on heating.** The lemon yellow powder is sparingly soluble in H<sub>2</sub>O and EtOH but soluble in aqueous NH<sub>3</sub>. It is obtained by heating H[Au(CN)<sub>2</sub>] at 110°. Wash it well with H<sub>2</sub>O and EtOH and dry it at 110°. It has an IR band at  $\nu_{\max}$  2239 cm<sup>-1</sup> typical for C≡N stretching vibration. [Glemser & Sauer *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1064 1965.] **CARE:** may evolve HCN.

**Gold (I) iodide** [10294-31-2] **M 323.9, m 120°(dec), d<sub>4</sub><sup>20</sup> 8.25.** It has been prepared by heating gold and iodine in a tube at 120° for 4 months. Since it decomposes to Au and I<sub>2</sub> in the presence of UV light and heat, then the main impurity is Au. The salt is therefore purified by heating at 120° with I<sub>2</sub> for several weeks. The crystals should be kept dry and in a cool place in the dark. [Weiss & Weiss *Z Naturforsch* **11B** 604 1956.]

**Gold (III) oxide hydrate** [1303-58-8] **M 441.9 + xH<sub>2</sub>O, evolves O<sub>2</sub> at 110°, pK<sub>1</sub><sup>25</sup> <11.7, pK<sub>2</sub><sup>25</sup> 13.36, pK<sub>3</sub><sup>25</sup> >15.3 [for Au(OH)<sub>3</sub>].** The most probable impurities are SO<sub>4</sub><sup>2-</sup> and Cl<sup>-</sup> ions. Dissolve it in strong boiling KOH solution (*ca* 5M) and precipitate (**care**) with excess of 3N H<sub>2</sub>SO<sub>4</sub>. Then shake and centrifuge, resuspend in H<sub>2</sub>O and repeat the washing several times until free from SO<sub>4</sub> and Cl ions. This gives a *wet* oxide which is dried in air, but decomposes to free gold in sunlight. It is advisable to keep it wet as it decomposes on drying (analyse wet sample). Store it away from light in the presence of H<sub>2</sub>O vapour. It evolves O<sub>2</sub> at 110°. It is insoluble in H<sub>2</sub>O but soluble in HCl and conc HNO<sub>3</sub>. [Roseveare & Buehrer *J Am Chem Soc* **49** 1221 1927.]

**Graphite (black lead, mineral carbon, plumbago)** [7782-42-5] **m 3652-3697°, d<sup>25</sup> 2.09, 2.23.** This carbon allotrope is obtained by mining (e.g. in Shrilanka, Canada). It is normally soft, forms black/gray scales, and although even small crystals of it are rare it is composed of crystallised carbon containing traces of SiO<sub>2</sub>, Fe and other minerals. Treat graphite with hot 1:1 HCl. Then filter, wash and the dried powdered is heated in an evacuated quartz tube at 1000° until a high vacuum is obtained. Cool this and store it in an atmosphere of helium [Craig et al. *J Phys Chem* **60** 1225 1956]. See also entry in Chapter 8.

**Helium** [7440-59-7] **M 4.0, m -272.2°/26atm, b -268.9°, d<sup>-270.3</sup> 0.147.** Dry the gas by passing it through a column of Linde 5A molecular sieves and CaSO<sub>4</sub>, then through an activated-charcoal trap cooled in liquid N<sub>2</sub>, to adsorb N<sub>2</sub>, argon, xenon and krypton. Also pass it over CuO pellets at 300° to remove hydrogen and hydrocarbons, over Ca chips at 600° to remove oxygen, and then over titanium chips at 700° to remove N<sub>2</sub> [Arnold & Smith *J Chem Soc, Faraday Trans 2* **77** 861 1981]. Its solubility in 100ml of H<sub>2</sub>O is 0.94ml at 25°, 1.05ml at 50° and 1.21ml at 75°.

**Hexachloroplatinic acid hydrate (H<sub>2</sub>PtCl<sub>6</sub>, chloroplatinic acid, platinum IV chloride solution)** [16941-12-1] **M 409.8 + H<sub>2</sub>O, m 60° (deliquescent solid).** If it is to be purified, or regenerated from Pt recovered from catalytic hydrogenations, it should be dissolved in *aqua regia* followed by evaporation to dryness and dissolution in the minimum volume of H<sub>2</sub>O. Then the aqueous solution is treated with saturated ammonium chloride until all the ammonium hexachloroplatinate [16919-58-7] separates. The (NH<sub>4</sub>)<sub>2</sub>PtCl<sub>6</sub> is filtered off and dried at 100°. Igniting this salt gives **Pt sponge**; dissolve the Pt sponge in *aqua regia*, boil to dryness, dissolve the residue in concentrated HCl, boil to dryness again and repeat the process. Protect it from light. [Wickers *J Am Chem Soc* **43** 1268 1921, Adams et al. *Org Synth Coll Vol I* 463, 466 1941, Bruce *J Am Chem Soc* **58** 687 1936.]

**Hexammine cobalt(III) chloride** [10534-89-1] **M 267.5.** It crystallises from warm water (8ml/g) on cooling.

[Bjerrum & McReynolds *Inorg Synth* **II** 217 1946.]

**Hexammine ruthenium(III) chloride** [14282-91-8] **M 309.6**. Crystallise it twice from 1M HCl.

**Hexarhodium hexadecacarbonyl** [28407-51-4] **M 1065.6, m 220°(dec, in air),  $d_4^{20}$  2.87**. It slowly loses CO when heated in air, but may be regenerated by heating at 80-200° in the presence of CO at 200 atmospheres pressure for 15 hours, preferably in the presence of Cu. It forms black crystals which are insoluble in hexane. It has bands at 2073, 2026 and 1800  $\text{cm}^{-1}$  in the IR. [Hieber & Lagally *Z Anorg Allgem Chem* **251** 96 1963, Corey & Dahl *J Am Chem Soc* **85** 1202 1963, Doyle et al. *Tetrahedron Lett* **22** 1783 1981.] **POISONOUS**.

**Hydrazine monohydrate ( $\text{N}_2\text{H}_4 \cdot \text{H}_2\text{O}$ )** [7803-57-8] **M 50.1, m 198°, b 118-122°/atm,  $d_4^{20}$  1.03**. It is best obtained by heating hydrazine sulphate (200g), NaOH (160g) and  $\text{H}_2\text{O}$  (75ml, exothermic) in a copper flask under reflux for 1.5 hours then distilled off (using a flame to remove all the hydrazine). The distillate (175ml) is a clear liquid which contains ~40-45% of  $\text{N}_2\text{H}_4$ . Note that hydrazine attacks glass, rubber and cork, and stainless steel equipment should be used. The percentage of hydrazine is determined by titration with standard acid (methyl orange indicator) or against standard iodine (starch indicator). *Hydrazine monohydrate* should contain 64% of  $\text{N}_2\text{H}_4$ . The ~40-45% solution may be concentrated by mixing it (144ml) with xylene (230ml) and distilling it through an efficient fractionating column (e.g. Hempel column). All the xylene passes over with about 85ml of  $\text{H}_2\text{O}$ . On distilling the residue, hydrated hydrazine (50ml) is obtained containing 80-85% of  $\text{N}_2\text{H}_4$ . This can be diluted with conductivity  $\text{H}_2\text{O}$  to 64%  $\text{N}_2\text{H}_4$  to give the *monohydrate*. Hydrazine and its hydrates have **VERY IRRITATING** and **TOXIC** vapours and should be used in an efficient fume cupboard. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** pp 469-472 1963.]

**Hydrazine hydrate ( $\text{N}_2\text{H}_4 \cdot x \text{H}_2\text{O}$ )** [10217-52-4] **M 32.05 + x 18.02**. Hydrated hydrazine can be obtained as above and diluted as required. Solutions containing various amounts of  $\text{H}_2\text{O}$  are available commercially.

**Hydrazine (anhydrous)** [302-01-2] **M 32.1, m 1.5-2.0°, b 47°/26mm, 56°/71mm, 113-113.5°/atm, n 1.470,  $d_4^{25}$  1.91,  $\text{pK}_1^{25}$  -0.88,  $\text{pK}_2^{25}$  8.11**. Hydrazine hydrate is dried by refluxing with an equal weight of KOH pellets for 3 hours, then distilled from fresh solid NaOH or BaO in a current of dry  $\text{N}_2$ . Use stainless steel or copper equipment. Hydrazine and its hydrates have **VERY IRRITATING** and **TOXIC** vapours and should be used in an efficient fume cupboard. Store in a well-stoppered vessel, preferably under  $\text{N}_2$ . It is a useful reducing agent. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** pp 469-472 1963.]

**Hydrazine dihydrochloride** [5341-61-7] **M 105.0, m 198°,  $d_4^{20}$  1.42**. It is recrystallised from aqueous EtOH and dried under vacuum over  $\text{CaSO}_4$ .

**Hydrazine monohydrochloride** [2644-70-4] **M 68.5, m 89°**. Prepare it by dropwise addition of cold conc HCl to cold liquid hydrazine in equimolar amounts. The crystals are harvested from water and are twice recrystallised from absolute MeOH and dried under a vacuum. [Kovack et al. *J Am Chem Soc* **107** 7360 1985.]

**Hydrazine sulfate** [10034-93-2] **M 130.1, m 254°**. Crystallise it from  $\text{H}_2\text{O}$ . Its solubility in  $\text{H}_2\text{O}$  is 3% at room temperature, but is very soluble in hot  $\text{H}_2\text{O}$ . It is a suspected **carcinogen**. [Adams & Brown *Org Synth Coll Vol I* 309 1941, Audrieth & Nickles *Inorg Synth I* 90 1939.]

**Hydrazoic acid (hydrogen azide, triazoic acid)** [7782-79-8] **M 43.0, m -80°, b 37°/atm,  $\text{pK}^{25}$  4.72**. The free acid is **HIGHLY EXPLOSIVE** and **POISONOUS**, so it is prepared and used in solution, and in an **efficient fume cupboard**. It is a very useful reagent for Schmidt and related reactions [Wolff *Organic Reactions* **3** 307 1946]. Solutions in  $\text{C}_6\text{H}_6$  or  $\text{CHCl}_3$  are prepared by making a paste from  $\text{NaN}_3$  (65g, 1mole) in warm  $\text{H}_2\text{O}$  (65ml), and added to  $\text{C}_6\text{H}_6$  or  $\text{CHCl}_3$  (400ml), cooled to 0°, and concentrated  $\text{H}_2\text{SO}_4$  (26.6mls, 0.5mol) is dropped into the stirred mixture while carefully controlling the temperature between 0° and 5°. The organic layer is separated, dried ( $\text{Na}_2\text{SO}_4$ ), and stored in a cold room. Its concentration is determined by shaking a small

aliquot with ten times its volume of H<sub>2</sub>O in a glass-stoppered flask, and titrating it against standard alkali. It can then be diluted to the desired concentration. [Audrieth & Gibbs *Inorg Synth* **1** 77 1939, Frost et al. *J Am Chem Soc* **55** 3516 1933, Schenk in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** p 472-474 1963.] The pure acid is a mobile, pungent explosive liquid, and **poisonous**, causing irritation of membranes, headaches, palpitation, lowering blood pressure and ataxia — *handle with great care*.

**Hydriodic acid** [10034-85-2] **M 127.9, b 127°(aqueous azeotrope), d<sub>4</sub><sup>20</sup> 1.701, pK<sup>25</sup> -8.56**. Iodine can be removed from aqueous HI, probably as the amine hydrogen triiodide, by three successive extractions using a 4% solution of Amberlite LA-2 (a long-chain aliphatic amine) in CCl<sub>4</sub>, toluene or petroleum ether (10ml per 100ml of acid). [Davidson & Jameson *Chem Ind (London)* 1686 1963.] Extraction with tributyl phosphate in CHCl<sub>3</sub> or other organic solvents is also suitable. *Alternatively*, a De-acidite FF anion-exchange resin column in the OH<sup>-</sup> form using 2M NaOH, then into its I<sup>-</sup> form by passing dilute KI solution through, can be used. Passage of an HI solution under CO<sub>2</sub> through such a column removes polyiodide. The column can be regenerated with NaOH. [Irving & Wilson *Chem Ind (London)* 653 1964]. The earlier method was to reflux with red phosphorus and distil in a stream of N<sub>2</sub>. The colourless product is stored in ampoules in the dark [Bradbury *J Am Chem Soc* **74** 2709 1952, Heisig & Frykholm *Inorg Synth* **I** 157 1939]. It fumes in moist air. **HARMFUL VAPOURS**.

**Hydrobromic acid** [10035-10-6] **M 80.9, b 125°(aqueous azeotrope, 47.5% HBr)/atm, d<sub>4</sub><sup>20</sup> 1.38 (34% HBr), pK<sup>25</sup> -8.69**. A solution of aqueous HBr *ca* 48% (w/w, constant boiling) is purified by distilling twice with a little red phosphorus, and the middle half of the distillate is taken. (The azeotrope at 760mm contains 47.8% (w/w) HBr.) [Hetzer et al. *J Phys Chem* **66** 1423 1962]. Free bromine can be removed by Irvine and Wilson's method for HI (see above), except that the column is regenerated by washing with an ethanolic solution of aniline or styrene. Hydrobromic acid can also be purified by aerating with H<sub>2</sub>S, distilling and collecting the fraction boiling at 125-127°. [Heisig & Andur *Inorg Synth* **I** 155 1939.] **HARMFUL VAPOURS**.

**Hydrochloric acid (muriatic acid)** [7647-01-0] **M 36.5, b 108.6°(aqueous azeotrope, 20.2% HCl), d<sub>4</sub><sup>20</sup> 1.09(20%), pK<sup>25</sup> -6.1**. It is readily purified by fractional distillation as the constant boiling point acid, following dilution with H<sub>2</sub>O. The constant-boiling fraction contains 1 mole of HCl in the following weights of distillate at the stated pressures: 179.555g (730mm), 179.766g (740mm), 179.979 (750mm), 180.193 (760mm), 180.407 (770mm). [Foulk & Hollingsworth *J Am Chem Soc* **45** 1220 1923.] **HARMFUL VAPOURS**.

**Hydrofluoric acid** [7664-39-3] **M 20.0, b 112.2°(aqueous azeotrope, 38.2% HF), d<sub>4</sub><sup>20</sup> 1.15 (47-53% HF), pK<sup>25</sup> 3.21**. It is freed from lead (Pb *ca* 0.002ppm) by co-precipitation with SrF<sub>2</sub>, by addition of 10ml of 10% SrCl<sub>2</sub> solution per kilogram of the concentrated acid. After the precipitate has settled, the supernatant is decanted through a filter in a hard-rubber or paraffin lined-glass vessel [Rosenqvist *Am J Sci* **240** 358 1942]. Pure aqueous HF solutions (up to 25M) can be prepared by isothermal distillation in polyethylene, polypropylene or platinum apparatus [Kwestroo & Visser *Analyst* **90** 297 1965]. It attacks glass and is used for etching glass. **HIGHLY TOXIC**.

**Hydrogen** [1333-74-0] **M 2.0, m -259.1°, b -252.9°, d 0.0889g/L (gas), 0.070g/L (liquid)**. It is usually purified by passing through a suitable absorption train of tubes. Carbon dioxide is removed with KOH pellets, soda-lime or NaOH pellets. Oxygen is removed with a "De-oxo" unit or by passage over Cu heated to 450-500° and Cu on Kieselguhr at 250°. Passage over a mixture of MnO<sub>2</sub> and CuO (Hopcalite) oxidises any CO to CO<sub>2</sub> (which is removed as above). Hydrogen can be dried by passage through dried silica-alumina at -195°, through a dry-ice trap followed by a liquid-N<sub>2</sub> trap packed with glass wool, through CaCl<sub>2</sub> tubes, or through Mg(ClO<sub>4</sub>)<sub>2</sub> or P<sub>2</sub>O<sub>5</sub>. Other purification steps include passage through a hot palladium thimble [Masson *J Am Chem Soc* **74** 4731 1952], through an activated-charcoal trap at -195°, and through a non-absorbent cotton-wool filter or small glass spheres coated with a thin layer of silicone grease. *Potentially VERY EXPLOSIVE in air or with O<sub>2</sub>*.

**Hydrogen bromide (anhydrous)** [10035-10-6] **M 80.9, b -66.8°/atm**. Dry it by passing it through Mg(ClO<sub>4</sub>)<sub>2</sub> towers. This procedure is **hazardous** [Stoss & Zimmermann *Ind Eng Chem* **17** 70 1939]. *Alternatively*, shake it with mercury, distil it through a -78° trap and condense it at -195°/10<sup>-5</sup>mm. It fumes in moist air. **HARMFUL VAPOURS**. It is soluble in H<sub>2</sub>O. A constant boiling aqueous solution of HBr has **b 126°/760mm**, and its HBr concentration is 47.4% (see hydrobromic acid above). It is soluble in AcOH. [Schneider & Johnson *Inorg Synth*

I 152 1939, Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 282-286 1963.]

**Hydrogen chloride** [7647-01-0] M 36.5, b -85°/760mm. Pass it through conc H<sub>2</sub>SO<sub>4</sub>, then over activated charcoal and silica gel. It fumes in moist air. Hydrogen chloride in gas cylinders contains ethylene, 1,1-dichloroethane and ethyl chloride. The latter two may be removed by fractionating the HCl through a trap cooled to -112°. Ethylene is difficult to remove. HCl fumes in moist air. **HARMFUL VAPOURS.** Its solubility in H<sub>2</sub>O is 82% at 0°. A constant boiling aqueous solution (azeotrope) has b 108.6°/760mm with an HCl concentration of ~20%, and is called *Hydrochloric acid* (*muriatic acid*) (see above). [Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 280-282 1963.]

**Hydrogen cyanide (anhydrous)** [74-90-8] M 27.0, b 25.7°/760mm, pK<sup>25</sup> 9.21 (aqueous acid). HCN is prepared from NaCN and H<sub>2</sub>SO<sub>4</sub>, and dried by passage through H<sub>2</sub>SO<sub>4</sub> and over CaCl<sub>2</sub>, then distilled in a vacuum system and degassed at 77°K before use [Arnold & Smith *J Chem Soc, Faraday Trans 2* 77 861 1981]. Cylinder HCN may contain stabilisers against explosive polymerisation, together with small amounts of H<sub>3</sub>PO<sub>4</sub>, H<sub>2</sub>SO<sub>4</sub>, SO<sub>2</sub>, and water. It can be purified by distillation over P<sub>2</sub>O<sub>5</sub>, then frozen in Pyrex bottles at Dry-ice temperature for storage. [Zeigler *Org Synth Coll Vol I* 314 1941, Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 658-660 1963.] Liquid HCN, like liquid ammonia, evaporates very slowly since the latent heat of evaporation is high and keeps it in the liquid state because the temperature of the liquid is lowered to below its boiling point. **EXTREMELY POISONOUS; all due precautions should be taken.**

**Hydrogen fluoride (anhydrous)** [7664-39-3] M 20.0, b 19.4°. It can be purified by trap-to-trap distillation, followed by drying over CoF<sub>2</sub> at room temperature and further distillation. *Alternatively*, it can be absorbed on NaF to form NaHF<sub>2</sub> which is then heated under vacuum at 150° to remove volatile impurities. The HF is regenerated by heating at 300° and is stored with CoF<sub>3</sub> in a nickel vessel, being distilled as required. (Water content should be *ca* 0.01%.) To avoid contact with base metal, use can be made of nickel, polychlorotrifluoroethylene and gold-lined fittings [Hyman et al. *J Am Chem Soc* 79 3668 1957]. An aqueous solution is *hydrofluoric acid* (see above). It is **HIGHLY TOXIC and attacks glass.**

**Hydrogen iodide (anhydrous)** [10034-85-2] M 127.9, b -35.5°. After removal of free iodine from aqueous HI, the solution is frozen, then covered with P<sub>2</sub>O<sub>5</sub> and allowed to melt under vacuum. The gas evolved is dried by passing through P<sub>2</sub>O<sub>5</sub> on glass wool. It can be freed from iodine contamination by repeated fractional distillation at low temperatures. It fumes in moist air, and an aqueous solution is *hydriodic acid* (see above). **HARMFUL VAPOURS.**

**Hydrogen peroxide** [7722-84-1] M 34.0, d<sub>4</sub><sup>20</sup> 1.110, pK<sup>25</sup> 11.65. The 30% material has been steam distilled using distilled water. Gross and Taylor [*J Am Chem Soc* 72 2075 1950] made 90% H<sub>2</sub>O<sub>2</sub> approximately 0.001M in NaOH and then distilled it under its own vapour pressure, keeping the temperature below 40°, the receiver being cooled with a Dry-ice/isopropyl alcohol slush. The 98% material has been rendered anhydrous by repeated fractional crystallisation in all-quartz vessels. **EXPLOSIVE IN CONTACT WITH ORGANIC MATERIAL.**

**Hydrogen peroxide urea adduct (UHP, urea hydrogen peroxide 1:1 complex, carbamide peroxide, Debrox, Hyperol)** [124-43-6] M 94.1, m. 85-90°(dec), 90°(dec). It is a safe alternative to H<sub>2</sub>O<sub>2</sub> in various oxidation reactions. It is commercially available in tablets ("rapidly soluble", equivalent to ~30% H<sub>2</sub>O<sub>2</sub>) or as a white powder (with 15-17% active oxygen). It is usually used without purification after assaying for active oxygen. This is done by titration with potassium permanganate or by iodometry, i.e. titration of liberated iodine when glacial acetic acid containing Fe<sup>3+</sup> and NaI are added. It can be recrystallised from 30% H<sub>2</sub>O<sub>2</sub> in a molar ratio of ~2:3 by heating in a pyrex dish for a few minutes at ~60°, cooling and allowed to crystallise slowly by evaporation in a crystallising dish. It forms elongated white needles, but if the solution is seeded just before crystallisation and shaken gently for a few seconds, then small plates are formed. Preferably collect the crystals by centrifugation at low temperature and dry them at 0° *in vacuo*. When dry, it is stable at room temperature and

it has been reported that the available oxygen content had not decreased noticeably after 12 months. However, it is best to store it dry at low temperature. It is soluble in organic solvents e.g. EtOH, Et<sub>2</sub>O, CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub> and Me<sub>2</sub>CO with slow decomposition, and its solubility in H<sub>2</sub>O is 40% where it also decomposes slowly. It decomposes slowly at 40-60°/20mm and at 55-70°/760mm in air, but decomposition appears to accelerate above 80°. It is very useful (and in many cases superior to *p*-chloroperbenzoic acid) in the oxidation of alkenes, (epoxides), aromatic hydrocarbons (to phenols), ketones (Baeyer-Villiger), sulfides (to sulfones) and N-heterocycles (to N-oxides) when using 5 to 10 molar ratios of oxidant in the presence of acetic or trifluoroacetic anhydrides. Care should be used with this reagent as it is **potentially explosive**. [Lu et al. *J Am Chem Soc* **63** 1508 1941, Cooper et al. *Synlett* 533 1990, *Beilstein* **3** H 54, **3** I 25, **3** II 45, **3** III 105, **3** IV 102.]

**Hydrogen sulfide** [7783-06-4] **M 34.1, b -59.6°, pK<sub>1</sub><sup>25</sup> 7.05, pK<sub>2</sub><sup>25</sup> 12.89**. Wash it, then pass the gas through a train of tubes containing saturated Ba(OH)<sub>2</sub> (2x), water (2x), and dilute HCl [Goates et al. *J Am Chem Soc* **73** 707 1951]. It is available in gas cylinders. **HIGHLY POISONOUS**.

**Hydroxylamine** [7803-49-8] **M 33.0, m 33.1°, b 56.5°/22mm, d<sub>4</sub><sup>20</sup> 1.226, pK<sup>20</sup> 5.96 (7.97)**. Crystallise it from *n*-butanol at -10°, collect it by vacuum filtration and wash it with cold diethyl ether. **Harmful vapours**. [Hurd *Inorg Synth* **I** 87 1939, Semon in *Org Synth Coll Vol* **I** 318 1932.]

**Hydroxylamine hydrochloride** [5470-11-1] **M 69.5, m 151°**. Crystallise the salt from aqueous 75% ethanol or boiling methanol, and dry it under vacuum over CaSO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub>. It has also been dissolved in a minimum of water and saturated with HCl; after three such crystallisations, it is dried under a vacuum over CaCl<sub>2</sub> and NaOH. Its solubility at 20° is 85% in H<sub>2</sub>O, 6% in EtOH and 12% in MeOH. [Hurd *Inorg Synth* **I** 87 1939, Semon in *Org Synth Coll Vol* **I** 318 1941.]

**Hydroxylamine sulfate** [10039-54-0] **M 164.1, m 170°(dec)**. Crystallise it from boiling water (1.6ml/g) by cooling to 0°.

**Hydroxylamine-O-sulfonic acid** [2950-43-8] **M 113.1, m 210-211°, 215°(dec), pK<sup>45</sup> 1.48**. Stir the solid vigorously with anhydrous Et<sub>2</sub>O and filter it off using large volumes of dry Et<sub>2</sub>O. Drain dry at the pump for 5 minutes and then for 12-14 hours in a vacuum. Store it in a vacuum desiccator over conc H<sub>2</sub>SO<sub>4</sub>. Determine the purity by oxidation of iodide to I<sub>2</sub>. It must be stored in a dry atmosphere at 0-4°. It decomposes slowly in H<sub>2</sub>O at 25° and more rapidly above this temperature. [Matsuguma & Andrieth *Inorg Synth* **V** 122 1957.]

**Hydroxyurea** [127-07-1] **M 76.1, m 70-72° (unstable form), m 133-136°, 141° (stable form), pK<sup>25</sup> 10.6**. Recrystallise hydroxyurea from absolute EtOH (10g in 150ml). Note that the rate of solution in boiling EtOH is slow (15-30 minutes). It should be stored in a cool dry place, but some decomposition could occur after several weeks. [Deghenghi *Org Synth Coll Vol* **V** 645 1973.] It is very soluble in H<sub>2</sub>O and can be crystallised from Et<sub>2</sub>O. [Kfod *Acta Chem Scand* **10** 256 1956, *Beilstein* **3** IV 170.]

**Hypophosphorous acid (Phosphinic acid)** [6303-21-5] **M 66.0, m 26.5°, d<sub>4</sub><sup>30</sup> 1.217, 1.13 and 1.04 for 50, 30-32, and 10% aqueous solutions resp, pK<sup>25</sup> 1.31 (H<sub>3</sub>PO<sub>2</sub>)**. Phosphorous acid is a common contaminant of commercial 50% hypophosphorous acid. Jenkins and Jones [*J Am Chem Soc* **74** 1353 1952] purified this material by evaporating about 600ml in a 1L flask at 40°, under reduced pressure (in N<sub>2</sub>), to a volume of about 300ml. After the solution was cooled, it was transferred to a wide-mouthed Erlenmeyer flask which was stoppered and left in a Dry-ice/acetone bath for several hours to freeze (if necessary, with scratching of the wall). When the flask was then left at *ca* 5° for 12 hours, about 30-40% of it liquefied, and was again filtered. This process was repeated, then the solid was stored over Mg(ClO<sub>4</sub>)<sub>2</sub> in a vacuum desiccator in the cold. Subsequent crystallisations from *n*-butanol by dissolving it at room temperature and then cooling in an ice-salt bath at -20° did not appear to purify it further. The free acid forms *deliquescent* crystals **m** 26.5° and is soluble in H<sub>2</sub>O and EtOH. The NaH<sub>2</sub>PO<sub>2</sub> salt can be purified through an anion exchange resin [Klement *Z Anorg Allgem Chem* **260** 267 1949.]

**Indium** [7440-74-6] **M 114.8, m 156.6°, b 2000°, d<sub>4</sub><sup>20</sup> 7.31.** Before use, the metal surface is cleaned with dilute HNO<sub>3</sub>, followed by thorough washing with water and an alcohol rinse. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 856 1963.]

**Indium (III) chloride** [10025-82-8] **M 211.2, m 586°, d<sub>4</sub><sup>20</sup> 4.0, pK<sub>1</sub><sup>25</sup> 3.54, pK<sub>2</sub><sup>25</sup> 4.28, pK<sub>3</sub><sup>25</sup> 5.16 (for aqueous In<sup>3+</sup>).** The anhydrous salt forms yellow deliquescent crystals which can be sublimed at 600° in the presence of Cl<sub>2</sub>/N<sub>2</sub> (1:1) [does not melt]. It is resublimed in the presence of Cl<sub>2</sub>/N<sub>2</sub> (1:10) and finally heated to 150° to expel excess Cl<sub>2</sub>. It is soluble in H<sub>2</sub>O and should be stored in a tightly closed container. [Baxter & Alter *J Am Chem Soc* **55** 1943 1933.]

**Indium (III) oxide** [1312-43-2] **M 277.6, d<sub>4</sub><sup>20</sup> 7.18, m sublimes at 850°.** Wash it with H<sub>2</sub>O and dry it below 850°. It volatilises at 850° and dissolves in hot mineral acids to form salts. Store it away from light because it darkens due to the formation of free In.

**Indium sulfate** [13464-82-9] **M 517.8.** Crystallise it from dilute aqueous H<sub>2</sub>SO<sub>4</sub>. It is *hygroscopic*; store it in a well-stoppered vessel.

**Indium (III) sulfate pentahydrate** [17069-79-3] **M 607.9, d<sub>4</sub><sup>20</sup> 3.44.** Dissolve the salt in strong H<sub>2</sub>SO<sub>4</sub> and slowly evaporate at *ca* 50°. Wash the crystals with glacial AcOH and then heat them in a furnace at a temperature of 450-500° for 6 hours. Its solubility in H<sub>2</sub>O is 5%. The *pentahydrate* is converted to an *anhydrous hygroscopic* powder on heating at 500° for 6 hours; but heating above this temperature over N<sub>2</sub> yields the oxide-sulfate. Evaporation of neutral aqueous solutions provides basic sulfates. [Baxter & Alter *J Am Chem Soc* **55** 1943 1933, Hattox & Vries *J Am Chem Soc* **58** 2126 1936.]

**Iodic acid** [7782-68-5] **M 175.9, m 118°(dec), d<sub>4</sub><sup>20</sup> 4.628, pK<sup>25</sup> 0.79.** Dissolve iodic acid in the minimum volume of hot dilute HNO<sub>3</sub>, filter and evaporate in a vacuum desiccator until crystals are formed. Collect the crystals and wash them with a little cold H<sub>2</sub>O and dry them in air in the dark. It is soluble in H<sub>2</sub>O: 269g/100ml at 20° and 295g/100ml at 40°. It is soluble in dilute EtOH and darkens on exposure to light. It is converted to HIO<sub>3</sub>.I<sub>2</sub>O<sub>5</sub> on heating at 70°, but at 220° complete conversion to HIO<sub>3</sub> occurs. [Lamb et al. *J Am Chem Soc* **42** 1636 1920, Bray & Caulkins *J Am Chem Soc* **53** 44 1931.]

**Iodine** [7553-56-2] **M 253.8, m 113.6°.** It is usually purified by vacuum sublimation. Preliminary purifications include grinding with 25% by weight of KI, blending with 10% BaO and subliming, subliming with CaO, grinding to a powder and treating with successive portions of H<sub>2</sub>O to remove dissolved salts, then drying, and recrystallising from \*benzene. Barrer and Wasilewski [*Trans Faraday Soc* **57** 1140 1961] dissolved I<sub>2</sub> in concentrated KI and distilled it, then steam distilled it three times and washed it with distilled H<sub>2</sub>O. Organic material is removed by sublimation in a current of O<sub>2</sub> over platinum at about 700°, the iodine being finally sublimed under vacuum. **HARMFUL VAPOURS.**

**Iodine monobromide** [7789-33-5] **M 206.8, m 42°.** The brown-black crystals are purified by repeated fractional crystallisation from its melt. The vapour dissociates on heating [Yost et al. *J Amer Chem Soc* **55** 552 1933, Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 291-292 1963].

**Iodine monochloride** [7790-99-0] **M 162.4, m 27.2°(α-form), 13.9°(β-form).** Purify it by repeated fractional crystallisation from its melt at low temperatures. The black crystals melt to a red-brown liquid. [Cornøg & Karges *Inorg Synth* **I** 165 1939.]

**Iodine pentafluoride** [7783-66-6] **M 221.9, m -8.0°, b 97°.** Rogers et al. [*J Am Chem Soc* **76** 4843 1954] removed dissolved iodine from IF<sub>5</sub> by agitating with a mixture of dry air and ClF<sub>3</sub> in a Fluoroethene beaker using a magnetic stirrer. The mixture is transferred to a still, and the more volatile impurities are pumped off as the pressure is reduced below 40mm. The still is gradually heated (kept at 40mm) to remove the ClF<sub>3</sub> before IF<sub>5</sub> distilled. Stevens [*J Org Chem* **26** 3451 1961] pumped IF<sub>5</sub> under vacuum from its cylinder, trapping it at -78°, then allowing it to melt in a stream of dry N<sub>2</sub>. **HARMFUL VAPOURS.** [Kwasnik in *Handbook of Preparative*

*Inorganic Chemistry (Ed. Brauer) Academic Press Vol I pp 159-160 1963.]*

**Iodine trichloride** [865-44-1] **M 233.3, m 33°, b 77°(dec).** Purify  $\text{ICl}_3$  by sublimation at room temperature.  
**Irritant vapours.** [Booth & Morris *Inorg Synth I* 167 1939.]

**Iridium** [7439-88-5] **M 192.2, m 2450°, b ~4500°,  $d_4^{20}$  22.65.** Iridium is a silver white hard solid which oxidises on the surface in air. Scrape the outer tarnished layer until silver clear and store it under paraffin. It is stable to acids but dissolves in aqua regia. [Gilchrist *Chem Rev* 32 277 1943.]

**Iridium (IV) chloride hydrate (hexachloroiridic acid)** [16941-92-7 ( $6\text{H}_2\text{O}$ ), 207399-11-9 ( $x\text{H}_2\text{O}$ )] **M 334.0+ $\text{H}_2\text{O}$ .** If it contains nitrogen, then repeatedly concentrate a conc HCl solution until free from nitrogen, and dry free from HCl in a vacuum over CaO until crystals are formed. The olive-green solid (yellow at  $\sim 700^\circ$ ) is very *hygroscopic*. [Woo & Yost *J Am Chem Soc* 53 884 1931, Grube in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer) Academic Press Vol II* p1592 1965.]

**Iron (wire)** [7439-89-6] **M 55.9, m 1535°.** Clean it in conc HCl, rinse in de-ionised water, then reagent grade acetone and dry it under vacuum.

**Iron enneacarbonyl (di-iron nonacarbonyl)** [15321-51-4] **M 363.7, m 100°(dec).** Wash it with EtOH and  $\text{Et}_2\text{O}$ , then dry it in air. It sublimes at  $35^\circ$  in a high vacuum. It forms dark yellow plates which are stable for several days when kept in small amounts. Large amounts, especially when placed in a desiccator, spontaneously *ignite* in a period of one day. It decomposes in moist air. It is insoluble in hydrocarbon solvents but forms complexes with several organic compounds. [Sheline & Pitzer *J Am Chem Soc* 72 1107 1950, Speyer & Wolf *Chem Ber* 60 1424 1927.] **TOXIC.**

**Iron pentacarbonyl (pentacarbonyl iron)** [13463-40-6] **M 195.9, m  $-20^\circ$ , b  $102.8^\circ/749\text{mm}$ ,  $103^\circ/760\text{mm}$ , n 1.520, d 1.490.** It is a pale yellow viscous liquid that is **PYROPHORIC** and readily absorbed by the skin. **HIGHLY TOXIC (protect from light and air).** It should be purified in a vacuum line by distilling and collecting in a trap at  $-96^\circ$  (toluene-Dry-ice slush). It has been distilled at atmospheric pressure (use a very efficient fume cupboard). At  $180^\circ$ /atmospheric pressure it decomposes to give Fe and CO. In UV light in petroleum ether it forms  $\text{Fe}_2(\text{CO})_9$  (see previous entry). [Hagen et al. *Inorg Chem* 17 1369 1978, Ewens et al. *Trans Faraday Soc* 35 6811 1939.]

**Lanthanum** [7439-91-0] **M 138.9, m  $920^\circ$ , b  $3470^\circ$ ,  $d_4^{20}$  6.16.** It is a shiny metal that slowly tarnishes in air due to oxidation. It slowly decomposes by  $\text{H}_2\text{O}$  in the cold and more rapidly on heating to form the hydroxide. The metal is cleaned by scraping off the tarnished areas until the shiny metal is revealed and stored under oil or paraffin. It burns in air at  $450^\circ$ . It exists in three forms:  $\alpha$ -form,  $\beta$ -form and  $\gamma$ -form with transition temperatures of  $310^\circ$  and  $864^\circ$ , respectively. [Spedding et al. *Ind Eng Chem* 44 553 1952.]

**Lead (II) bromide** [10031-22-8] **M 367.0, m  $373^\circ$ .** Crystallise it from water containing a few drops of HBr (25ml of water per gram  $\text{PbBr}_2$ ) between  $100^\circ$  and  $0^\circ$ . A neutral solution is evaporated at  $110^\circ$ , and the crystals that separate are collected by rapid filtration at  $70^\circ$  and dried at  $105^\circ$  (to give the *monohydrate*). Its solubility in  $\text{H}_2\text{O}$  is 0.5% (at  $\sim 10^\circ$ ) and 5% (at  $\sim 100^\circ$ ). To prepare the *anhydrous* bromide, the hydrate is heated for several hours at  $170^\circ$  and then in a Pt boat at  $200^\circ$  in a stream of HBr and  $\text{H}_2$ . Finally it is fused [Clayton et al. *J Chem Soc, Faraday Trans 1* 76 2362 1980]. **POISONOUS.**

**Lead (II) chloride** [7758-95-4] **M 278.1, m  $501^\circ$ .** Crystallise it from distilled water at  $100^\circ$  (33ml/g) after filtering through sintered-glass and adding a few drops of HCl, by cooling. After three crystallisations the solid is dried under vacuum or under anhydrous HCl vapour by heating slowly to  $400^\circ$ . The solubility in  $\text{H}_2\text{O}$  is 0.07% at  $\sim 10^\circ$ , and 0.43% at  $\sim 100^\circ$ . **POISONOUS.**

**Lead (II) iodide** [10101-63-0] **M 461.0, m  $402^\circ$ .** It crystallises from a large volume of water. The solubility in

H<sub>2</sub>O is 1.1% at ~10°, and 3.3% at ~100°. **POISONOUS.**

**Lead monoxide** [1317-36-8] **M 223.2, m 886°.** Higher oxides are removed by heating under vacuum at 550° with subsequent cooling under vacuum. It is red at room temperature but becomes yellow at high temperatures (~480°) reversibly. [Ray & Ogg *J Am Chem Soc* **78** 5994 1956, Kwestroo et al. *J Inorg Nucl Chem* **29** 39 1967.] **POISONOUS.**

**Lead nitrate** [10099-74-8] **M 331.2, m 470°.** Precipitate it twice from a hot (60°) concentrated aqueous solution by adding HNO<sub>3</sub>. The precipitate is sucked dry on a sintered-glass funnel, then transferred to a crystallising dish which is covered by a clock glass and left in an electric oven at 110° for several hours [Beck et al. *Trans Faraday Soc* **55** 331 1959]. After two recrystallisations of ACS grade, no metals above 0.001ppm were detected. **POISONOUS.**

**Lithium (metal)** [7439-93-2] **M 6.9, m 180.5°, b 1342°, d<sub>4</sub><sup>20</sup> 0.534.** After washing with petroleum ether to remove storage oil, lithium is fused at 400° and then forced through a 10-micron stainless-steel filter with argon pressure. It is again melted in a dry-box, skimmed, and poured into an iron distillation pot. After heating under a vacuum to 500°, cooling and returning it to the dry-box for a further cleaning of its surface, the lithium is distilled at 600° using an all-iron distillation apparatus [Gunn & Green *J Am Chem Soc* **80** 4782 1958].

**Lithium aluminium hydride** [16853-85-3] **M 37.9, m 125°(dec).** Extract it with Et<sub>2</sub>O, and, after filtering, the solvent is removed under vacuum. The residue is dried at 60° for 3 hours, under high vacuum [Ruff *J Am Chem Soc* **83** 1788 1961]. It is a strong reducing agent. Store it in aliquots in a strictly dry atmosphere, and use in these aliquot quantities. **It IGNITES in the presence of a small amount of water and reacts with it EXPLOSIVELY.** [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 805 1963.]

**Lithium amide** [7782-89-0] **M 23.0, m 380-400°, d<sub>4</sub><sup>17.5</sup> 1.178.** Purify it by heating at 400° while NH<sub>3</sub> is passed over it in the upper of two crucibles (the upper crucible is perforated). The LiNH<sub>2</sub> will drip into the lower crucible through the holes in the upper crucible. The product is cooled in a stream of NH<sub>3</sub>. Protect it from air and moisture, store it under N<sub>2</sub> in a clear glass bottle sealed with paraffin. Store it in small quantities so that all the material is used once the bottle is opened. If the colour of the amide is yellow, it should be destroyed as it is likely to have oxidised and to **EXPLODE**. On heating above 450° it is decomposed to Li<sub>2</sub>NH, which is stable up to 750-800°. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 463 1963, Greenlee & Henne *Inorg Synth* **II** 135 1953.]

**Lithium azide** [19597-69-4] **M 49.0, m 115-298° (dec), d<sub>4</sub><sup>20</sup> 1.008, pK<sup>25</sup> 4.72 (HN<sub>3</sub>).** Digest ~1g with 15ml of 96% EtOH at 35°, filter and dry it in air at temperatures below 80°. Store it in a cool place and treat it as potentially explosive. Its solubility in H<sub>2</sub>O is 66.4% at 16°, and 20.26% at 16° in EtOH. [Hofmann-Bang *Acta Chem Scand* **11** 581 1957, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 581 1963.]

**Lithium borohydride** [16949-15-8] **M 21.8, m 268°, b 380°(dec), d<sub>4</sub><sup>20</sup> 0.66.** It is crystallised from Et<sub>2</sub>O, and pumped free of ether at 90-100° during 2 hours [Schaeffer et al. *J Am Chem Soc* **78** 729 1956]. Store it dry as it decomposes slowly in moist air. [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 775 1963.]

**Lithium bromide** [7550-35-8] **M 86.8, m 550°.** Crystallise it several times from water or EtOH, then dry it under high vacuum for 2 days at room temperature, followed by drying at 100°. Its solubility in H<sub>2</sub>O is 167% at ~20°, and 250% at ~100°. It is *deliquescent* and should be stored in a tightly stoppered vessel.

**Lithium carbonate** [554-13-2] **M 73.9, m 552°, 618°.** Crystallise it from water. Its solubility decreases as the temperature is raised. The solubility in H<sub>2</sub>O is 1.3% at ~10°, and 0.7% at ~100°. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 987 1963, Caley & Elving *Inorg Synth* **I** 1 1939.]

**Lithium chloride** [7447-41-8] **M 42.4, m 600°, 723°**. Crystallise it from water (1ml/g) or MeOH and dry it for several hours at 130°. Other metal ions can be removed by preliminary crystallisation from hot aqueous 0.01M disodium EDTA. It has also been crystallised from conc HCl, fused in an atmosphere of dry HCl gas, cooled under dry N<sub>2</sub> and pulverised in a dry-box. Kolthoff and Bruckenstein [*J Am Chem Soc* **74** 2529 1952] precipitated it with ammonium carbonate, washed it with Li<sub>2</sub>CO<sub>3</sub> five times by decantation and finally with suction, then dissolved it in HCl. The LiCl solution is evaporated slowly with continuous stirring in a large evaporating dish, the dry powder being stored (while still hot) in a desiccator over CaCl<sub>2</sub>.

**Lithium fluoride** [7789-24-4] **M 25.9, m 842°, 848°, b 1676°, 1681°, d<sub>4</sub><sup>20</sup> 2.640**. Possible impurities are LiCO<sub>3</sub>, H<sub>2</sub>O and HF. These can be removed by calcining it at red heat, then pulverizing it with a Pt pestle and storing it in a paraffin bottle. Its solubility in H<sub>2</sub>O is 0.27% at 18°. It volatilises between 1100-1200°. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 235 1963].

**Lithium hydride**. [7580-67-8] **M 7.95, m 680°, d<sub>4</sub><sup>20</sup> 0.76-0.77**. It should be a white powder; otherwise replace it. It darkens rapidly on exposure to air and is decomposed by H<sub>2</sub>O to give H<sub>2</sub> and LiOH, and reacts with lower alcohols. One gram in H<sub>2</sub>O liberates 2.8L of H<sub>2</sub> (could be explosive). [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 987 1963.]

**Lithium hydroxide monohydrate** [1310-66-3 (H<sub>2</sub>O), 1310-65-2 (anhydrous)] **M 42.0, m 471°, d 1.51, pK<sup>25</sup> 13.82**. It crystallises from water (3ml/g) as the *monohydrate*. It *dehydrates* at 150° in a stream of CO<sub>2</sub>-free air, and sublimes at 220° with some decomposition [Cohen *Inorg Synth* **V** 3 1957, Bravo *Inorg Synth* **VII** 1 1963].

**Lithium iodate** [13765-03-2] **M 181.9**. Crystallise it from water and dry it in a vacuum oven at 60°.

**Lithium iodide** [10377-51-2] **M 133.8, m 73° (3H<sub>2</sub>O), 469°, b 1171°, d<sub>4</sub><sup>20</sup> 4.06**. Crystallise it from hot water (0.5ml/g) by cooling in a CaCl<sub>2</sub>-ice EtOH or from an acetone- Dry-ice bath. Dry it under a vacuum over P<sub>2</sub>O<sub>5</sub> for 1 hour at 60° and then at 120°. It is *deliquescent* and should be stored in dark tightly stoppered vessels.

**Lithium nitrate** [7790-69-4] **M 68.9, m 253°, d<sub>4</sub><sup>20</sup> 2.38**. It crystallises from water or EtOH. Dry it at 180° for several days by repeated melting under vacuum. If it is crystallised from water keeping the temperature above 70°, formation of *trihydrate* is avoided. The *anhydrous* salt is dried at 120° and stored in a vacuum desiccator over CaSO<sub>4</sub>. After the 99% pure salt was recrystallised 3 times, it contained: metal (ppm) Ca (1.6), K (1.1), Mo (0.4), Na (2.2). [Donnan & Burt *J Chem Soc* **83** 335 1903.]

**Lithium nitrite monohydrate** [13568-33-7] **M 71.0**. Crystallise it from water by cooling from room temperature.

**Lithium perchlorate** [7791-03-9] **M 106.4, m 236°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Crystallise it from water or 50% aqueous MeOH. It is rendered *anhydrous* by heating the *trihydrate* at 170-180° in an air oven. It can then be recrystallised twice from acetonitrile and again dried under vacuum [Mohammad & Kosower *J Am Chem Soc* **93** 2713 1971]. **SKIN IRRITANT**.

**Lithium sulfate (anhydrous)** [10377-48-7] **M 109.9, loses H<sub>2</sub>O at 130° and m 859°, d<sub>D</sub><sup>20</sup> 2.21**. Crystallise it from H<sub>2</sub>O (4ml/g) by partial evaporation, and dry it above 130° *in vacuo*.

**Lithium tetrafluoroborate** [14283-07-9] **M 93.7, pK<sup>25</sup> 13.82 (Li<sup>+</sup>), pK<sup>25</sup> -4.9 (for HBF<sub>4</sub>)**. Dissolve it in THF just below its solubility, filter from insoluble material and evaporate it to dryness in a vacuum below 50°. Wash the residue with dry Et<sub>2</sub>O, and pass dry N<sub>2</sub> gas over the solid and finally heat it in an oven at 80-90°. Its solubility in Et<sub>2</sub>O is 1.3g in 100ml at 25°; in THF it is 71g in 100ml at 25°. It is *hygroscopic* and is an **IRRITANT**. [Elliott et al. *J Am Chem Soc* **74** 5211 1952, **75** 1753 1953.]

**Lithium thiocyanate (lithium rhodanide)** [556-65-0] **M 65.0, pK<sup>25</sup> -1.85 (for HSCN)**. It crystallises from H<sub>2</sub>O as the *dihydrate*, but on drying at 38-42° it gives the *monohydrate*. It can be purified by allowing an aqueous solution to crystallise in a vacuum over P<sub>2</sub>O<sub>5</sub>. The crystals are collected, dried out *in vacuo* at 80°/P<sub>2</sub>O<sub>5</sub>

in a stream of pure  $N_2$  at  $110^\circ$ . [Coates & Taylor *J Chem Soc* 1245 1936.]

**Magnesium** [7439-95-4] **M 24.3, m 651°**, **b 1100°**, **d<sub>4</sub><sup>20</sup> 1.739**. It slowly oxidises in moist air and tarnishes. If dark in colour, do not use. The shiny solid should be degreased by washing with dry  $Et_2O$ , dry it *in vacuo* and keep it in a  $N_2$  atmosphere. It can be activated by stirring it in  $Et_2O$  containing a crystal of  $I_2$  then filtering it off, before drying and storing. [Gmelin's Magnesium (8<sup>th</sup> edn) **27A** 121 1937.]

**Magnesium bromide (anhydrous)** [7789-48-2] **M 184.1, m 711°**, **d<sub>4</sub><sup>20</sup> 3.72**. Crystallise it from  $EtOH$  or  $H_2O$  (3.3g/ml). Dry it in a vacuum at  $\sim 150^\circ$ , or heat the *hydrate* in a stream of  $HCl$ . It is very *deliquescent*. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 909 1963.]

**Magnesium chloride hexahydrate** [7791-18-6] **M 203.3, m  $\sim 100^\circ$ (dec)**, **pK<sub>1</sub><sup>25</sup> 10.3, pK<sub>2</sub><sup>25</sup> 12.2 (for  $Mg^{2+}$  hydrolysis)**. Crystallise it from hot water (3.3g/ml) by cooling. Dry it in a vacuum at  $\sim 175^\circ$ , or heat the *hydrate* in a stream of  $HCl$ . When the hydrate is heated above  $180^\circ$  it is hydrolysed to the oxychloride ( $Mg_2OCl_2$ ). It is *deliquescent*; store it in a well-stoppered vessel. [Bryce-Smith *Inorg Synth* **VI** 9 1960, Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 908 1963.]

**Magnesium iodate tetrahydrate** [7790-32-1] **M 446.2**. Crystallise from water (0.2g/ml) between  $100^\circ$  and  $0^\circ$ .

**Magnesium iodide** [10377-58-9] **M 278.1, m 634°**. Crystallise it from water (0.8g/ml) by partial evaporation in a desiccator. It is *deliquescent* and should be stored in the dark. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 910 1963.]

**Magnesium nitrate hexahydrate** [13446-18-9] **M 256.4, m  $\sim 95^\circ$ (dec)**. Crystallise the nitrate from water (2.5mlg) by partial evaporation in a desiccator. It is *deliquescent* and is soluble in  $EtOH$ . After two recrystallisations, ACS grade salt has: metal (ppm) Ca (6.2), Fe (8.4), K (2), Mo (0.6), Na (0.8), Se (0.02).

**Magnesium perchlorate (Anhydrone, Dehydrite)** [10034-81-8 (anhydrous)] **M 259.2, m  $>250^\circ$** , **pK<sup>25</sup> -2.4 to -3.1 (for  $HClO_4$ )**. Crystallise it from water to give the *hexahydrate* **M 331.3** [13346-19-0]. Coll et al. [*J Am Chem Soc* **81** 1284 1959] removed traces of unspecified contaminants by washing it with small portions of  $Et_2O$  and drying in a vacuum (**CARE**). The anhydrous salt is commercially available as an ACS reagent, and is as efficient a dehydrating agent as  $P_2O_5$  and is known as "Dehydrite" or "Anhydrone". [Smith et al. *J Am Chem Soc* **44** 2255 1922 and *Ind Eng Chem* **16** 20 1924.] It is *hygroscopic*; keep it in a tightly closed container. It is **EXPLOSIVE in contact with organic materials, and is a SKIN IRRITANT**.

**Magnesium sulfate (anhydrous)** [7487-88-9] **M 120.4, m 1127°**. Crystallise it from warm  $H_2O$  (1g/ml) by cooling. Dry the *heptahydrate* (*Epsom salt*) at  $\sim 250^\circ$  until it loses 25% of its weight. Its solubility in  $H_2O$  is 36% at  $20^\circ$ , 55% at  $60^\circ$  and 74% at  $100^\circ$ ; above  $110^\circ$  the solubility decreases with rise of temperature. Store it in a sealed container.

**Manganese decacarbonyl  $Mn_2(CO)_{10}$**  [10170-69-1] **M 390.0, m 151-152°**, **154-155°(sealed tube)**, **d<sup>25</sup> 1.75**. Golden yellow crystals which in the absence of  $CO$  begin to decompose at  $110^\circ$ , and on further heating yield a metallic mirror. In the presence of 3000psi of  $CO$  it does not decompose on heating to  $250^\circ$ . It is soluble in common organic solvents, insoluble in  $H_2O$ , not very stable in air, to heat or UV light. It dissolves in a lot of  $*C_6H_6$  and can be crystallised from it. It distils with steam at  $92-100^\circ$ . It can be purified by sublimation under reduced pressure ( $<0.5mm$ ) at room temperature to give well-formed golden yellow crystals. If the sample is orange coloured, this sublimation leads to a mixture of golden-yellow and dark red crystals of the carbonyl and carbonyl iodide, respectively, which can be separated by hand picking under a microscope. Separation followed by resublimation provides the pure compounds. **POISONOUS**. [Brimm et al. *J Am Chem Soc* **76** 3831 1954, Closson et al. *J Am Chem Soc* **80** 6167 1958, **82** 1325 1960.]

**Manganous bromide (anhydrous)** [13446-03-2] **M 214.8, m 695° (4H<sub>2</sub>O)** [10031-20-6] **M 286.8, m 64°(dec)**. It forms rose-red *deliquescent* crystals which are soluble in  $EtOH$ . The  $H_2O$  is removed by heating at

100° then in HBr gas at 725°, or dry it in an atmosphere of N<sub>2</sub> at 200°.

**Manganous chloride tetrahydrate** [13446-34-9, 7773-01-5 (anhydrous)] **M 197.9, m 58°, 87.5°, 650° (anhydrous), b 1190° (anhydrous), d<sub>4</sub><sup>20</sup> 2.01.** It crystallises from water (0.3ml/g) on cooling. The red-rose *tetrahydrate* melts at ~52° and forms the *dihydrate salt* which loses all its H<sub>2</sub>O at 198° to give MnCl<sub>2</sub>. It is soluble in EtOH.

**Manganous sulfate monohydrate** [10034-96-5 (H<sub>2</sub>O), 15244-36-7 (xH<sub>2</sub>O)] **M 169.0, d<sub>4</sub><sup>20</sup> 2.75.** Crystallise it from water (0.9ml/g) at 54-55° by evaporating about two-thirds of the water. It *dehydrates* above 400°.

**Mercuric bromide** [7789-47-1] **M 360.4, m 238.1°, b 320°, d<sub>4</sub><sup>20</sup> 5.73.** Crystallise it from hot saturated ethanolic solution, dry and keep it at 100° for several hours under a vacuum, then sublime it. [Garrett *J Am Chem Soc* **61** 2744 1939.] Its solubility in H<sub>2</sub>O is 0.6% at 20°, and 22% at 100°; in EtOH it is 30% at 25°; and in MeOH it is 69.6% at 25°. [Wagenknecht & Juza *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1109 1965.] **POISONOUS.**

**Mercuric chloride** [7487-94-7] **M 271.5, m 276°, b 304°, d<sub>4</sub><sup>20</sup> 5.6.** Crystallise it twice from distilled water, dry it at 70° and sublime it under high vacuum. Its solubility in H<sub>2</sub>O is 4.3% at ~0°, 6.6% at ~10° and 54% at ~100°. It is soluble in EtOH and is extracted into Et<sub>2</sub>O from an aqueous solution. It is **very POISONOUS** and 0.2-0.4g is fatal. The antidote is immediate administration of white of egg as an emetic.

**Mercuric cyanide** [592-04-1] **M 252.6, m 320°(dec), d<sub>4</sub><sup>20</sup> 4.00.** Crystallise it from water. The solubility in H<sub>2</sub>O is 8% at ~20° and 33% at ~100°; in EtOH it is 8% at ~20° and in MeOH it is 25% at ~20°. [Blitz *Z Anorg Allgem Chem* **170** 161 1928.] **POISONOUS.**

**Mercuric iodide (red)** [7774-29-0] **M 454.4, m 259°(yellow >130°), b ~350°(subl), d<sub>4</sub><sup>20</sup> 6.3.** Crystallise it from MeOH or EtOH and wash it repeatedly with distilled water (solubility is 0.006% at ~25°). It has also been mixed thoroughly with excess 0.001M iodine solution, filtered, washed with cold distilled water, rinsed with EtOH and Et<sub>2</sub>O, and dried in air. It changes colour reversibly to yellow at ~130°. [Friend *Nature* **109** 341 1922.] **POISONOUS.**

**Mercuric oxide (yellow)** [21908-53-2] **M 216.6, m 500°(dec).** Dissolve it in HClO<sub>4</sub> and precipitate it with NaOH solution. It is yellow when cold and changes to red at ~130° reversibly. **POISONOUS.**

**Mercuric thiocyanate** [592-85-8] **M 316.8, m 165°(dec), pK<sup>25</sup> -1.85 (for HSCN).** Recrystallise it from H<sub>2</sub>O, and it can give various crystal forms depending on conditions. Its solubility in H<sub>2</sub>O is 0.069% at 25°, but is more soluble at higher temperatures. It decomposes to Hg above 165°. **POISONOUS.** [Mason & Forgeng *J Phys Chem* **35** 1121 1931, Birckenbach & Kolb *Chem Ber* **68** 919 1935.]

**Mercurous nitrate dihydrate** [7782-86-7 (2H<sub>2</sub>O), 7783-34-8 (H<sub>2</sub>O), 10415-75-5 (anhydrous)] **M 561.2, m 70°(dec), d<sub>4</sub><sup>20</sup> 4.78, pK<sup>25</sup> 2.68 (for Hg<sub>2</sub><sup>2+</sup> hydrolysis).** Its solubility in H<sub>2</sub>O containing 1% HNO<sub>3</sub> is 7.7%. Recrystallise it from a warm saturated solution of dilute HNO<sub>3</sub> and cool to room temperature slowly to give elongated prisms. Rapid cooling gives plates. The colourless crystals should be stored in the dark. **POISONOUS.** [Grdenic *J Chem Soc* 1312 1956.]

**Mercurous sulfate** [7783-36-0] **M 497.3, d<sub>4</sub><sup>20</sup> 7.56.** The white-yellow powder is recrystallised from dilute H<sub>2</sub>SO<sub>4</sub>, dried in a vacuum under N<sub>2</sub>, and stored in the dark. Its solubility in H<sub>2</sub>O is 0.6% at 25°. It is hydrolysed by excessive washing with H<sub>2</sub>O to form the greenish-yellow *basic salt* Hg<sub>2</sub>SO<sub>4</sub>.Hg<sub>2</sub>O.H<sub>2</sub>O. **POISONOUS.**

**Mercury** [7439-97-6] **M 200.6, m -38.9°, b 126°/1mm, 184°/10mm, 261°/100mm, 356.9°/atm, d<sub>4</sub><sup>20</sup> 13.534.** After air has been bubbled through mercury for several hours to oxidise metallic impurities, it is filtered to remove coarser particles of oxide and dirt, then sprayed through a 4-ft column containing 10% HNO<sub>3</sub>. It is washed with distilled water, dried with filter paper and distilled under vacuum. [Schenk in *Handbook of*

*Preparative Inorganic Chemistry (Ed. Brauer) Academic Press Vol I p 8 1963.]*

**Molybdenum** [7439-98-7] **M 95.9, m 2622°, b ~4825°, d<sup>25</sup> 10.3g/ml, resistivity of 5.0 μΩ-cm at 20°.** The dark-gray metal is commercially available in varying degrees of purity from 98 to 99.99%, as foil with thickness from 0.025mm (5.8g 150 x 150mm) to 1.0mm (25.6g, 50mm x 50mm or 102.4g, 100mm x 100 mm), or black powder from <150μm to 1-2μm particle size as well in nanopowder form of <100nm (BET) size. It is also available as wire of 1.0mm in diameter. The metal is stable in air but is oxidised at red heat to MoO<sub>3</sub>; is unreactive towards dilute acids or alkalis but reacts with strong HNO<sub>3</sub>, H<sub>2</sub>SO<sub>4</sub>, fused KClO<sub>3</sub> or KNO<sub>3</sub>, F<sub>2</sub> (at ~ 25°), Cl<sub>2</sub> and Br<sub>2</sub> (at red heat). [**CARE:** it is potentially quite toxic—particularly the nano sized particles.]

**Molybdenum (VI) dichloride dioxide (MoO<sub>2</sub>Cl<sub>2</sub>)** [13637-68-8] **M 198.8, d<sup>25</sup> 6.31g/ml.** It is prepared from MoO<sub>2</sub> (40.0g, 313mmol) by placing it in a test tube with a wide side arm and a central tube which ends ~2cm above the surface of the oxide. The side arm is connected to a large round bottomed flask which vents through a bubbler containing H<sub>2</sub>SO<sub>4</sub>. The apparatus is flushed with argon, and the tube containing the oxide is immersed in an oil bath which is kept at 140° for 12 hours to dry the oxide. The temperature of the tube is then raised to 160° and a stream of Cl<sub>2</sub> gas (pre-dried by bubbling through two H<sub>2</sub>SO<sub>4</sub> traps) is allowed to flow over the hot oxide for 10 hours. The wide tube connecting the test tube and the flask is wrapped with a heating tape to encourage the MoO<sub>2</sub>Cl<sub>2</sub> to sublime into the flask as it is formed. Stirring the contents of the flask (magnetic stirrer bar) hinders the clogging entry and exit tubes. MoO<sub>2</sub>Cl<sub>2</sub> collects as ivory flakes that are quite pure for most purposes, but can be resublimed if necessary. All due **PRECAUTIONS** should be taken as chlorine is a **TOXIC GAS** and concentrated H<sub>2</sub>SO<sub>4</sub> is used in the bubblers. [Schrock et al. *J Am Chem Soc* **112** 3875 1990, Epperson et al. *Inorg Synth* **7** 168 1963.] Commercial fluffy MoO<sub>2</sub>Cl<sub>2</sub> gives poor yields of substitution products (e.g. with ArNHSiMe<sub>3</sub>), but the THF complex MoO<sub>2</sub>Cl<sub>2</sub>(THF)<sub>2</sub> [556907-19-8; 12081-12-8] is easier to handle, reacts in a similar way, and gives much higher yields of substitution products. It is readily prepared in solution by carefully adding solid MoO<sub>2</sub>Cl<sub>2</sub> to dry THF at -30°. Addition of THF to MoO<sub>2</sub>Cl<sub>2</sub> is too exothermic to keep under control. [Schrock et al. *J Am Chem Soc* **112** 3875 1990.] Krauss & Huber prepared it by adding MoO<sub>2</sub>Cl<sub>2</sub> (2.0g, 10mmol) in four portions to THF (7ml. 6.5g, 90mmol) while shaking. The yellow solution is diluted with light petroleum (5ml), filtered, the filtrate is diluted further with light petroleum (10ml) when the complex crystallises out. The supernatant is decanted off, the solid is washed several times with light petroleum (10ml lots), and dried at room temperature in a stream of N<sub>2</sub> to give analytically pure MoO<sub>2</sub>Cl<sub>2</sub>(THF)<sub>2</sub> (3.2g, 93%) as very pale yellow needles, **m > 50° (dec)**, that are soluble in THF, slightly soluble in Et<sub>2</sub>O but insoluble in light petroleum [Krauss & Huber *Chem Ber* **94** 2864 1961]. It is very useful for preparing a variety of Mo complexes and catalysts.

**Molybdenum hexacarbonyl** [13939-06-5] **M 264.0, m 150°(dec), b 156°.** Sublime it in a vacuum before use [Connor et al. *J Chem Soc, Dalton Trans* 511 1986]. **TOXIC.**

**Molybdenum hexafluoride** [7783-77-9] **M 209.9, m 17.5°, b 35°/760mm, d<sub>4</sub><sup>20</sup> 2.543.** Purify the hexafluoride by low-temperature trap-to-trap distillation over pre-dried NaF. It is *hygroscopic*, fumes in moist air and is hydrolysed readily by H<sub>2</sub>O. [Oppengard et al. *J Am Chem Soc* **82** 3825 1960, Anderson & Winfield *J Chem Soc, Dalton Trans* 337 1986, Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer) Academic Press Vol I p 259 1963.]* **Poisonous vapours.**

**Molybdenum (IV) oxide [molybdenum (IV) dioxide, MoO<sub>2</sub>]** [18868-43-4] **M 127.9, d<sup>25</sup> 6.47.** It is prepared by grinding together ~10-15g of a mixture in a 2:1 ratio of MoO<sub>3</sub> (dried by heating in at 500° for 1 hour) and Mo metal powder in a quartz combustion tube under argon, or in an evacuated tube (CARE when opening), at 800° for 70 hours. It forms a brown-violet powder, oblong needles or thick platelets with a metallic lustre. [Conroy & Ben-Dor *Inorg Synth* **14** 149 1973, **30** 105 1997, Herzog et al. in *Handbuch der Preparativen Anorganische Chemie (Ed. Brauer) Enke Verlag Stuttgart Vol 3 p 1542 1981.]* Beware of **TOXIC** vapours.

**Molybdenum trichloride** [13478-18-7] **M 202.3, m 1027°, d<sub>4</sub><sup>20</sup> 3.74.** Boil it with 12M HCl, wash it with absolute EtOH and dry it in a vacuum desiccator. It is a brown-red powder soluble in H<sub>2</sub>O, EtOH or Et<sub>2</sub>O and gives a blue solution in conc H<sub>2</sub>SO<sub>4</sub>. [Hein & Herzog *Handbook of Preparative Inorganic Chemistry (Ed. Brauer) Academic Press Vol II p 1404 1965.]*

**Molybdenum trioxide [molybdenum (VI) oxide, MoO<sub>3</sub>]** [1313-27-5] **M 143.9, m 795°, b 1155°, d<sub>4</sub><sup>20</sup> 4.60.** MoO<sub>3</sub> is prepared by adding HNO<sub>3</sub> to an aqueous solution of ammonium molybdate and stirring for several hours which precipitates H<sub>2</sub>MoO<sub>4</sub>. The acid is filtered off, washed with H<sub>2</sub>O, heated at 150° for 1-2 hours and the hydrated oxide is dried at 450°. It recrystallises in rhombs from water (1g/50ml) between 70° and 0°, and is dried in air at ~500° for 1 hour before use. Its solubility in H<sub>2</sub>O is 0.1% at 18°, and 2% at 70°. It is a white powder which turns yellow reversibly on heating. It sublimes readily at >780°/760mm in a quartz tube to give pure MoO<sub>3</sub>. [Hein & Herzog *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1412 1965, Herzog et al. in *Handbuch der Preparativen Anorganische Chemie* (Ed. Brauer) Enke Verlag Stuttgart Vol 3 p 1544 1981.] It is used in the synthesis of Mo compounds, and of the ternary reduced molybdenum oxide Pr<sub>4</sub>Mo<sub>9</sub>O<sub>18</sub>, which contains the previously unknown Mo<sub>7</sub>, Mo<sub>13</sub>, and Mo<sub>19</sub> clusters; and is a precursor for the preparation of fast ion conductors and superconductors [Arulraj *Chem Mater* 14 2492 2002, Lacorre *Nature* 404 856 2000]. The oxide forms hydrates and alcoholates which have different IR spectra [Krauss & Huber *Chem Ber* 94 2864 1961]. Beware of **TOXIC** vapours.

**Molybdic acid (H<sub>2</sub>MoO<sub>4</sub>)** [7782-91-4] **M 162.0, 180.0 (monohydrate), d<sup>25</sup> 3.124 (for monohydrate), pK<sub>1</sub><sup>25</sup> 0.9 (proton addition), pK<sub>2</sub><sup>20</sup> 4.00, pK<sub>3</sub><sup>20</sup> 4.21.** Treatment of an aqueous solution of 0.4N Na<sub>2</sub>MoO<sub>4</sub> [1L, or of (NH<sub>4</sub>)<sub>2</sub>MoO<sub>4</sub> as in the preceding entry] at 60° with 30% HNO<sub>3</sub> (1L, prepared from 300ml of concentrated acid, d 1.42, in 1L of H<sub>2</sub>O with cooling) and allowing to stand at 25° for several days deposited the theoretical amount of canary yellow monoclinic crystals which are collected, washed with ice water and dried *in vacuo* over H<sub>2</sub>SO<sub>4</sub> for 2 weeks to give H<sub>2</sub>MoO<sub>4</sub>·H<sub>2</sub>O. This is referred to as **α-molybdic acid** which on warming in H<sub>2</sub>O at 70° is transformed to the white anhydrous **β-molybdic acid**. These acids have different X-Ray spectra and vapour pressures. **Colloidal molybdic acid** is obtained by dialysing an aqueous solution of (NH<sub>4</sub>)<sub>2</sub>MoO<sub>4</sub> and HCl, and forms a gum on evaporation. By heating these acids, or the ammonium salts, provide white MoO<sub>3</sub> which on heating further at 500° in the presence of H<sub>2</sub> gives reddish-brown MoO<sub>2</sub>, and at 1200° a grey powder of metallic Mo is obtained (see preceding entries). Polymeric forms are known such as **tetramolybdic acid** [H<sub>2</sub>Mo<sub>4</sub>O<sub>16</sub>, pK<sub>1</sub> 1.4, pK<sub>2</sub> 1.2, Chauveau et al. *Bull Soc Chim Fr* 1190 1959] and **heptamolybdic acid** [H<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub>, pK<sub>5</sub><sup>25</sup> ~3.7, pK<sub>6</sub><sup>25</sup> 4.33, Sasaki et al *J Inorg Nucl Chem* 9 93 1959]. Note that in aqueous alkaline solution the main Mo(VI) acid species are MoO<sub>4</sub><sup>2-</sup>, but in neutral and acidic media polynuclear species exist [cf Sasaki & Sillén *Arkiv Kemi* 29 253 1968, and Sasaki et al *J Inorg Nucl Chem* 9 93 1959]. [Herzog et al. in *Handbuch der Preparativen Anorganische Chemie* (Ed. Brauer) Enke Verlag Stuttgart Vol 3 p 1544 1981, Rosenheim *Z Anorg Chem* 50 320 1906, Peters et al. *Z Anor Allgem Chem* 365 14 1969; and for pKs see also Sasaki & Sillén *Acta Chem Scand* 18 1014 1964, Rohwer & Cruywagen *J S African Chem Inst* 16 26 1963, Rohwer & Cruywagen *J S African Chem Inst* 17 145 1964.]

**Monocalcium phosphate dihydrate (monobasic)** [7789-77-7 (2H<sub>2</sub>O), 7757-93-9 (anhydrous)] **M 154.1, m 200°(dec, loses H<sub>2</sub>O at 100°), d<sub>4</sub><sup>20</sup> 2.2.** Crystallise it from a near-saturated solution in 50% aqueous reagent grade phosphoric acid at 100° by filtering through fritted glass and cooling to room temperature. The crystals are filtered off, and this process is repeated three times using fresh acid. For the final crystallisation the solution is cooled slowly with constant stirring to give thin plate crystals that are filtered off on a fritted glass funnel, washed free of acid with anhydrous acetone and dry in a vacuum desiccator [Egan et al. *J Am Chem Soc* 78 1811 1956].

**Neodymium chloride hexahydrate** [13477-89-9] **M 358.7, m 124°, pK<sub>1</sub><sup>25</sup> 8.43 (for Nd<sup>3+</sup> hydrolysis).** Neodymium chloride forms large purple prisms from concentrated solutions of dilute HCl. They are soluble in H<sub>2</sub>O (2.46 parts in 1 part of H<sub>2</sub>O) and EtOH, and lose H<sub>2</sub>O at 160°.

**Neodymium(II) iodide** [61393-36-0] **M 398.1.** This one electron reductant can be prepared in large quantities (~40g) by direct reaction of the metal and iodine at 600° [Evans et al. *Inorg Chem* 42 3097 2003]. The black solid NdI<sub>2</sub> can be kept at room temperature for months in the absence of solvent. Solutions can be stored under argon for several hours at -15° but should be used as soon as possible. The stability is considerably reduced if N<sub>2</sub> is used as inert atmosphere, instead of argon, even at -30°. In tetrahydrofuran under argon it promotes a pseudo-Barbier reaction, i.e. rapid reductive coupling of primary and secondary (but not tertiary) halides with ketones, e.g. butylchloride and cyclohexanone provide almost quantitative yield of 1-butylhexan-1-ol [Evans et

al. *Org Lett* **5** 2041 2003]. Thus behaving somewhat like a Grignard reagent. This solution also has Birch reduction-type reactivity, e.g. reduces naphthalene to 1,4-dihydronaphthalene.

**NdI<sub>2</sub> (THF)<sub>x</sub> solution** is prepared by transferring pre-cooled THF (45ml, at -15°) *via* a cannula into a septum-capped flask containing black NdI<sub>2</sub> (1.0g, 2.5mmol) and a magnetic stirrer bar at -15° whereby the colour of the solution becomes purple, and is stirred for 45 minutes. An aliquot should show the concentration to be 0.05M by complexometric titration [method of Evans & Allen *J Am Chem Soc* **122** 2118 2000]. This solution is used for reactions and should be performed under argon.

**Neodymium(III) iodide (anhydrous)** [13813-24-6] **M 524.7, m 775°, b 1370°**. It is an almost black crystalline powder (green when ground) that is soluble in hot and cold H<sub>2</sub>O. It can be prepared by the method of Bochkarev [Bochkarev & Fagin *Chem Eur J* **5** 2990 1999] where a small quantity of Nd metal is placed in a quartz crucible in a quartz reactor [see Evans et al. *Inorg Chem* **42** 3097 2003], heated to 600°, and I<sub>2</sub> is added, then small amounts of metal and I<sub>2</sub> are added alternately. Every addition of I<sub>2</sub> results in an orange glow in the mixture. When addition is complete (total ≥3 equivalents) the apparatus is cooled, and the crucible containing NdI<sub>3</sub> is transferred to an argon filled glovebox and the salt is ground (with a pestle and mortar) to a green powder. The metal content is analysed by titration (complexometric metal analysis is performed by dissolving the salt in H<sub>2</sub>O at ~25°, evaporating, ashing the residue at 500°, dissolving in HCl and the analysis is carried out in hexamethylenetetramine buffer with xylenol orange as indicator and EDTA as titrant: Schwarzenbach & Flaschka "Complexometric Titrations" Metheun, London p 194 1969; Evans et al. *J Am Chem Soc* **103** 6672 1981). It should contain at least 27.5% Nd (theoretical is 29.5% Nd). [Evans & Workman *Organometallics* **24** 1989 2005.]

**Potassium graphite KC<sub>8</sub> with NdI<sub>3</sub>** has also been used for reductive coupling and is prepared *in situ*. [Weitz & Rabinovitz *J Chem Soc, Perkin Trans I* 117 1993]. KC<sub>8</sub> can be prepared in a Schlenk line, or in a glove box, by adding K metal (0.541g, 13.8mmol) to a scintillating vial containing a Teflon stirbar then Graphite (1.242g, 12.9mmol), and the mixture is stirred and heated until a bronze-coloured powder results. The NdI<sub>3</sub>/KC<sub>8</sub>/THF is prepared by adding precooled (-15° *via* a cannula) THF into a septum-capped Schlenk flask (~25ml) containing NdI<sub>3</sub>, KC<sub>8</sub> and a Teflon stirbar thus producing a purple solution which is stirred at -15° for 40 minutes and is ready for the reaction. The alkyl halide is injected, stirred for 1 minute, and is followed by the aldehyde or ketone. [Evans & Workman *Organometallics* **24** 1989 2005.]

**Neodymium nitrate hexahydrate** [16454-60-7] **M 438.4, m 70-72°**. It crystallises with 5 and 6 molecules of H<sub>2</sub>O from concentrated solutions in dilute HNO<sub>3</sub> by slow evaporation; 1 part is soluble in 10 parts of H<sub>2</sub>O.

**Neodymium oxide** [1313-97-9] **M 336.5, m 2320°**. Dissolve it in HClO<sub>4</sub>, precipitate it as the oxalate with doubly recrystallised oxalic acid, wash it free of soluble impurities, dry it at room temperature and ignite it in a platinum crucible at higher than 850° in a stream of oxygen. It is a blue powder. [Tobias & Garrett *J Am Chem Soc* **80** 3532 1958.]

**Neon** [7440-01-9] **M 20.2**. Pass the gas through a copper coil packed with 60/80 mesh 13X molecular sieves which is cooled in liquid N<sub>2</sub>, or through a column of Ascarite (NaOH-coated silica/asbestos adsorbent).

**Nickel bromide** [13462-88-9] **M 218.5, m 963°(loses H<sub>2</sub>O at ~ 200°)**. Crystallise it from dilute HBr (0.5ml/g) by partial evaporation in a desiccator. The *anhydrous* salt is yellow, but the *trihydrate* is green.

**Nickel chloride hexahydrate** [7791-20-0 (6H<sub>2</sub>O), 69098-15-3 (xH<sub>2</sub>O), 7718-54-9 (*anhydrous*)] **M 237.7**. It crystallises from dilute HCl to form the green *hexahydrate*. At 70° this dehydrates to the *tetrahydrate*, and at higher temperatures it forms the *anhydrous* salt. It sublimes in yellow hexagonal scales in a stream of HCl. Store it in a desiccator as it is *deliquescent*. [Hart & Partington *J Chem Soc* 104 1943.]

**Nickel nitrate hexahydrate** [13478-00-7] **M 290.8, m 57°**. Crystallise it from water (3.3g/ml) by partial evaporation in a desiccator. Store it in a desiccator as it is *deliquescent*.

**Nickel(II) perchlorate hexahydrate** [Ni(ClO<sub>4</sub>)<sub>2</sub> · 6H<sub>2</sub>O] [13520-61-1] **M 365.7, m 140°, 200°, 209° (sealed tube), d<sup>0</sup> 1.570, d<sup>20</sup> 1.583, d<sup>40</sup> 1.597, d<sup>50</sup> 1.646, d<sup>60</sup> 1.597**. The greenish blue hexagonal prisms of this salt are

obtained by double decomposition between  $\text{NiSO}_4$  and  $\text{Ca}(\text{ClO}_4)_2$  in aqueous solution, filtering off the insoluble  $\text{BaSO}_4$ , and concentrating the filtrate. It recrystallises from  $\text{H}_2\text{O}$ , and drying the crystals in a vacuum desiccator over  $\text{H}_2\text{SO}_4$  or  $\text{CaCl}_2$  provides the *hexahydrate*. Alternatively, it is prepared from freshly precipitate  $\text{NiO}$  or  $\text{NiCO}_3$  and aqueous  $\text{HClO}_4$ , filtering off the excess of  $\text{NiCO}_3$  and concentrating the filtrate to crystallisation. The solubility of  $\text{Ni}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$  in  $\text{H}_2\text{O}$  (g/100ml) is 217 ( $0^\circ$ ), 236 ( $10^\circ$ ), 245 ( $20^\circ$ ), 267 ( $30^\circ$ ), 280 ( $35^\circ$ ), 273 ( $40^\circ$ ), 311 ( $50^\circ$ ), 295 ( $55^\circ$ ), 280 ( $60^\circ$ ). The salt form hydrates with 2, 4, 5, 6, 7, and 9 molecules of  $\text{H}_2\text{O}$ . The *hexahydrate* is normally obtained by crystallisation from  $\text{H}_2\text{O}$  and is the more stable hydrate. It forms the *tetrahydrate* at  $110^\circ/760\text{mm}/36$  hours or  $70^\circ/1\text{mm}/66$  hours, which yields the *dihydrate* at  $130^\circ/760\text{mm}/42$  hours or  $100^\circ/1\text{mm}/72$  hours. It should be kept in a desiccator over  $\text{CaCl}_2$  as it is quite *hygroscopic*. The UV in EtOH has  $\lambda_{\text{max}}$  ( $\log \epsilon$ ) at 395 ( $\sim 0.68$ ), 660 ( $\sim 0.22$ ), 720 ( $\sim 0.25$ ) nm. Evaporation of a solution in ether with dry air gives  $\text{Ni}(\text{ClO}_4)_2 \cdot \text{Et}_2\text{O}$  as a yellow powder, and with dioxan it forms  $\text{Ni}(\text{ClO}_4)_2 \cdot 2\text{C}_4\text{H}_8\text{O}_2$ . A solution in  $\text{Me}_2\text{CO}$  at  $\sim 2 \times 10^{-4}\text{M}$  is green with a characteristic absorption at 410nm [Katzin *Nature* **182** 1013 1958]; it complexes also with acetylacetone [Fernelius et al. *J Chem Phys* **59** 235 1955], ethylene, trimethylenediamine [Cotton & Harris *J Phys Chem* **59** 1203 1955], DMF [Pflaum, & Popov *Anal Chim Acta* **13** 167, 169 1955], and with pyridine it forms blue crystals of  $[\text{Ni}(\text{C}_5\text{H}_5\text{N})_6](\text{ClO}_4)_2 \cdot 4\text{H}_2\text{O}$  [Weinland et al. *Arch Pharm* **265** 352 1927],  $\text{Ni}(\text{ClO}_4)_2 \cdot 4(\text{C}_5\text{H}_5\text{N})$  and  $\text{Ni}(\text{ClO}_4)_2 \cdot 6(\text{C}_5\text{H}_5\text{N})$  [Sinha *J Indian Chem Soc* **35** 865 1958, Sinha & Ray *J Indian Chem Soc* **20** 32 1943]. [Salvadori *Gazzetta* **42** I 458 1912, Veeraiah & Qureshi *J Indian Chem Soc* **21** 127 1944, Bernath et al. *J Prakt Chem* **143** [2] 302 1935, Freund & Schneider *J Am Chem Soc* **81** 4780 1959, Gmelin's *Handbuch der Anorganischen Chemie, Nickel Teil B-Lieferung 21*, Verlag Chemie GMBH Weinheim, System 57, pp 596-601 1966.]

**Nickel sulfate hexahydrate** [10101-97-0] **M 262.9, m loses  $5\text{H}_2\text{O}$  at  $100^\circ$ , becomes anhydrous at  $\sim 280^\circ$ ,  $d^{25}$  2.07.** It crystallises from  $\text{H}_2\text{O}$  in bluish-green tetragonal crystals of the  $\alpha$ -form which undergo a transition at  $53.3^\circ$  to the  $\beta$ -form that has green transparent monoclinic crystals. The crystals are stable at  $\sim 40^\circ$ , become blue in colour and then opaque in air due to slow efflorescence. Its solubility in  $\text{H}_2\text{O}$  is 0.6/ml at  $0^\circ$  and 3.4g/ml at  $100^\circ$ . It is sparingly soluble in EtOH, slightly more soluble in MeOH, soluble in aqueous NaOH, and very soluble in aqueous ammonia to give the bluish-green monoclinic  $(\text{NH}_4)_2\text{Ni}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$  double salt. The pH of an aqueous solution is about 4.5, i.e. weakly acidic. (See also the heptahydrate below.)

**Nickel sulfate heptahydrate** [10101-98-1, 7786-81-4 (anhydrous)] **M 280.9, m loses  $5\text{H}_2\text{O}$  at  $100^\circ$ , anhydrous m at  $\sim 280^\circ$ ,  $d^{25}$  1.948.** The sulfate crystallises from warm water (4g/ml) or dilute  $\text{H}_2\text{SO}_4$  as bright green monoclinic crystals on cooling. It is isomorphous with Epsom salt (see  $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}$  above). Prolonged exposure to air gives the blue *tetrahydrate*. On heating above  $118^\circ$ , it is converted to the *dihydrate*, which at  $>280^\circ$  is converted to the *yellow anhydrous*  $\text{NiSO}_4$  which does not react with HCl.

**Nickel(II) sulfide** **M 90.8 (NiS, mellirite, bronze-yellow,  $d^{25}$  5.3-5.6), 240.2 ( $\text{Ni}_3\text{S}_2$  [12035-72-2] Heazelwoodite,  $d^{25}$  5.8) and ( $\text{Ni}_3\text{S}_2/\text{NiS}$  black powder [16812-54-7]  $d^{25}$  5.3-5.6), m  $790^\circ$ ,  $797^\circ$ .** The sulfide is generally obtained as a highly insoluble black precipitate by adding  $\text{NH}_4\text{OH}$  and  $\text{NH}_4\text{SH}$  to a nickel salt. It dissolves slowly in excess of  $(\text{NH}_4)_2\text{S}_x$  and re-precipitates on boiling, exposure to air or addition of acid. A black dense sulfide is also obtained by boiling a nickel salt with  $\text{Na}_2\text{S}_2\text{O}_3$ . It is insoluble in dilute HCl, slowly soluble in concentrated HCl, readily in  $\text{HNO}_3$  and aqua regia. The sulfides should be washed well with  $\text{H}_2\text{O}$ , dried, and analysed for Ni. Various forms have been prepared [Dunn & Rideal *J Chem Soc* **123** 1242 1923.]

**Niobium (Colombium) (V) chloride** [10026-12-7] **M 270.2, m  $204.7-209.5^\circ$ , b  $\sim 250^\circ$  (begins to sublime at  $125^\circ$ ),  $d_4^{20}$  2.75.** It forms yellow, very deliquescent crystals which decompose in moist air to liberate HCl. Keep it in a dry box flushed with  $\text{N}_2$  in the presence of  $\text{P}_2\text{O}_5$ . Wash it with  $\text{CCl}_4$  and dry it over  $\text{P}_2\text{O}_5$ . The yellow crystals can contain a few small, dirty white pellets among the yellow needles. These should be easily picked out. Upon grinding in a dry box, however, they turn yellow.  $\text{NbCl}_5$  has been sublimed and fractionated in an electric furnace. [Epperson *Inorg Synth* **VII** 163 1963, Alexander & Fairbrother *J Chem Soc* suppl 233 1949.]

**Nitric acid (Aqua fortis)** [7697-37-2] **M 63.0, m  $-42^\circ$ , b  $83^\circ$ ,  $d^{25}$  1.5027, [Constant boiling acid has composition w/w of 68%  $\text{HNO}_3$  + 32%  $\text{H}_2\text{O}$ , b  $120.5^\circ$ ,  $d_4^{20}$  1.41],  $\text{pK}^{25}$  -1.27 (1.19).** The acid is obtained

colourless (approx. 92%) by direct distillation of fuming  $\text{HNO}_3$  under reduced pressure at 40-50° with an air leak at the head of the fractionating column. **Concentrated nitric acid** is an aqueous solution containing 70-71% of  $\text{HNO}_3$  ( $d_4^{20}$  1.4134). Store it in a desiccator that is kept in a refrigerator away from light which causes the formation of  $\text{NO}_2$ . Nitrite-free  $\text{HNO}_3$  can be obtained by vacuum distillation from urea. [Ward et al. *Inorg Synth III* 13 1950, Kaplan & Schechter *Inorg Synth IV* 53 1953.] If “fuming nitric acid” (90%) is yellow in colour (due to the presence of oxides of nitrogen), then treat 100mls with urea (0.5g) and bubble dry air through it until it is colourless (~20 minutes) [Freeman & Shepard *Org Synth* 43 84 1932]. Acid that is free from oxides does not discolour drops of 1N  $\text{KMnO}_4$ . **Anhydrous  $\text{HNO}_3$**  (100%,  $d_4^{20}$  1.5129) is obtained on distilling a mixture of equal volumes of “fuming nitric acid” and concentrated  $\text{H}_2\text{SO}_4$  [Liang *Org Synth Coll Vol III* 804 1955]. It is an oxidising agent, its vapours irritate and cauterise tissues, and cause the skin to turn yellow in colour. The “fuming acid” is particularly dangerous, should be handled in an efficient fume cupboard, and is used more as an oxidant than for acidification — use EYE PROTECTION.

**Nitric oxide** [10102-43-9] **M 30.0, b -151.8°**. Bubble the gas through 10M  $\text{NaOH}$  which removes  $\text{NO}_2$ . It can also be freed from  $\text{NO}_2$  by passage through a column of Ascarite followed by a column of silica gel held at -197°K. The gas is dried with solid  $\text{NaOH}$  pellets or by passing through silica gel cooled at -78°, followed by fractional distillation from a liquid  $\text{N}_2$  trap. This purification does not eliminate nitrous oxide. Other gas scrubbers sometimes used include one containing conc  $\text{H}_2\text{SO}_4$  and another containing mercury. It is freed from traces of  $\text{N}_2$  by the freeze and thaw method. [Blanchard *Inorg Synth II* 126 1946, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 485-487 1963.] **TOXIC**.

**Nitrogen** [7727-37-9] **M 28.0, b -195.8°**. Cylinder  $\text{N}_2$  can be freed from oxygen by passage through **Fieser's solution** [which comprises 2g sodium anthraquinone-2-sulfonate and 15g sodium hydrosulfite dissolved in 100ml of 20%  $\text{KOH}$ ; see Fieser, *J Am Chem Soc* 46 2639 1924] followed by scrubbing with saturated lead acetate solution (to remove any  $\text{H}_2\text{S}$  generated by the Fieser solution), conc  $\text{H}_2\text{SO}_4$  (to remove moisture), then soda-lime (to remove any  $\text{H}_2\text{SO}_4$  and  $\text{CO}_2$ ). *Alternatively*, after passage through Fieser's solution,  $\text{N}_2$  can be dried by washing with a solution of the metal ketyl from benzophenone and  $\text{Na}$  wire in absolute diethyl ether. [If ether vapour in  $\text{N}_2$  is undesirable, the ketyl from liquid  $\text{Na-K}$  alloy under xylene can be used.]

Another method for removing  $\text{O}_2$  is to pass the nitrogen through a long, tightly packed column of  $\text{Cu}$  turnings, the surface of which is constantly renewed by scrubbing it with ammonia (sg 0.880) solution. The gas is then passed through a column packed with glass beads moistened with conc  $\text{H}_2\text{SO}_4$  (to remove ammonia), through a column of packed  $\text{KOH}$  pellets (to remove  $\text{H}_2\text{SO}_4$  and to dry the  $\text{N}_2$ ), and finally through a glass trap packed with chemically clean glass wool immersed in liquid  $\text{N}_2$ . Nitrogen has also been purified by passage over  $\text{Cu}$  wool at 723°K and  $\text{Cu(II)}$  oxide [prepared by heating  $\text{Cu(NO}_3)_2 \cdot 6\text{H}_2\text{O}$  at 903°K for 24 hours] and then into a cold trap at 77°K.

A typical dry purification method consists of a mercury bubbler (as trap), followed by a small column of silver and gold turnings to remove any mercury vapour, towers containing anhydrous  $\text{CaSO}_4$ , dry molecular sieves or  $\text{Mg(ClO}_4)_2$ , a tube filled with fine  $\text{Cu}$  turnings and heated to 400° by an electric furnace, a tower containing soda-lime, and finally a plug of glass wool as filter. Variations include tubes of silica gel, traps containing activated charcoal cooled in a Dry-ice bath, copper on Kieselguhr heated to 250°, and  $\text{Cu}$  and  $\text{Fe}$  filings at 400°. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 458-460 1963.]

**Nitrosyl chloride** [2696-92-6] **M 65.5, m -64.5°. b -5.5°**. It is an orange gas with a suffocating odour. It has been fractionally distilled at atmospheric pressure in an all-glass, low-temperature still, taking the fraction boiling at -4° and storing it in sealed tubes. *Alternatively*, the gas is dried by  $\text{CaCl}_2$  and passed through  $\text{H}_2\text{SO}_4$  when  $\text{Cl}_2$  passes on, but  $\text{NOCl}$  is absorbed to form **nitrososulfuric acid** ( $\text{NO} \cdot \text{HSO}_4$ ) which on warming with  $\text{NaCl}$  evolves pure  $\text{NOCl}$  [Tilden *J Chem Soc* 27 630 1874.] It is decomposed by  $\text{H}_2\text{O}$  and alkali, and forms compounds with metal chlorides e.g.  $\text{FeCl}_3 \cdot \text{NOCl}$ . [Coleman *Inorg Synth I* 55 1939.]

**Nitrous oxide** (laughing gas) [10024-97-2] **M 44.0, b -88.5°**. Wash the gas with concentrated alkaline pyrogallol solution, to remove  $\text{O}_2$ ,  $\text{CO}_2$ , and  $\text{NO}_2$ , then dry it by passing it through columns of  $\text{P}_2\text{O}_5$  or Drierite, and collecting in a dry trap cooled in liquid  $\text{N}_2$ . It is further purified by freeze-pump-thaw and distillation cycles

under vacuum [Ryan & Freeman *J Phys Chem* **81** 1455 1977, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 484-485 1963].

**Osmium tetroxide (osmic acid)** [20816-12-0] M 524.2, m 40.6°, b 59.4°/60mm, 71.5°/100mm, 109.3°/400mm, 130°/760mm,  $d_D^{20}$  5.10,  $pK_1^{25}$  7.2,  $pK_2^{25}$  12.2,  $pK_3^{25}$  13.95,  $pK_4^{25}$  14.17 (H<sub>4</sub>OsO<sub>6</sub>). It is **VERY TOXIC** and should be manipulated in a very efficient fume cupboard. It attacks the eyes severely (**use also eye and face protection**) and is a good oxidising agent. It is volatile and has a high vapour pressure (11mm) at room temperature. It sublimes and volatilises well below its boiling point. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, H<sub>2</sub>O (7.24% at 25°), CCl<sub>4</sub> (375% at 25°), EtOH and Et<sub>2</sub>O. It is estimated by dissolving a sample in a glass-stoppered flask containing 25ml of a solution of KI (previously saturated with CO<sub>2</sub>) and acidified with 0.35M HCl. After gentle shaking in the dark for 30 minutes, the solution is diluted to 200ml with distilled H<sub>2</sub>O saturated with CO<sub>2</sub> and titrated with standard thiosulfate using starch as indicator. This method is not as good as the gravimetric method. Hydrazine hydrochloride (0.1 to 0.3g) is dissolved in 3M HCl (10ml) in a glass-stoppered bottle. After warming to 55-65°, a weighed sample of OsO<sub>4</sub> solution is introduced, and the mixture is digested on a water bath for 1 hour. The mixture is transferred to a weighed glazed crucible and evaporated to dryness on a hot plate. A stream of H<sub>2</sub> is started through the crucible, and the crucible is heated over a burner for 20-30 minutes. The stream of H<sub>2</sub> is continued until the crucible is cooled to room temperature, and then the H<sub>2</sub> is displaced by CO<sub>2</sub> in order to avoid rapid combustion of H<sub>2</sub>. Finally the crucible is weighed. [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II pp 1603 1965, Anderson & Yost *J Am Chem Soc* **60** 1822 1938.]

§ Available commercially on a polymer support.

**Oxygen** [7782-44-7] M 32.00, m -218.4°, b -182.96°,  $d^{-183}$  1.149,  $d^{-252.5}$  1.426. Purify it by passing the gas over finely divided platinum at 673°K and Cu(II) oxide (see under nitrogen) at 973°, then condensed in a liquid N<sub>2</sub>-cooled trap. **HIGHLY EXPLOSIVE in contact with organic matter.**

**Palladium (II) chloride** [7647-10-1] M 177.3, m 678-680°. The *anhydrous* salt is insoluble in H<sub>2</sub>O and dissolves in HCl with difficulty. The *dihydrate* forms red *hygroscopic* crystals that are readily reduced to Pd. Dissolve it in conc HCl through which dry Cl<sub>2</sub> is bubbled. Filter this solution which contains H<sub>2</sub>PdCl<sub>4</sub> and H<sub>2</sub>PdCl<sub>6</sub> and on evaporation it yields a residue of pure PdCl<sub>2</sub>. [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed Brauer) Academic Press Vol II p 1582 1965, Mozingo *Org Synth Coll Vol III* 685 1955.] *Alternatively* (fume cupboard), Pd metal is dissolved in *aqua regia*, and converted to the chloride by repeated evaporation with concentrated HCl. The dark red palladous chloride is dried at 150-180° in a stream of dry chlorine gas to give the *anhydrous* salt. [Kharasch, Seyler and Mayo *J Am Chem Soc* **60** 882 1938.]

**Palladium (II) cyanide** [2035-66-7] M 158.1. The yellow solid should be washed well with H<sub>2</sub>O and dry in air. [Bigelow *Inorg Synth II* 245 1946]. **POISONOUS.**

**Perchloric acid** [7601-90-3] M 100.5,  $d_4^{20}$  1.665,  $pK^{25}$  -2.4 to -3.1 (HClO<sub>4</sub>). The 72% acid has been purified by double distillation from silver oxide under vacuum: this frees the acid from metal contamination. Distillation at atmospheric pressure is **dangerous and explosive**. The *anhydrous* acid is obtained by adding gradually 400-500ml of oleum (20% fuming H<sub>2</sub>SO<sub>4</sub>) to 100-120ml of 72% HClO<sub>4</sub> in a reaction flask cooled in an ice-bath. The pressure is reduced to 1mm (or less), with the reaction mixture at 20-25°. The temperature is gradually raised during 2 hours to 85°; the distillate is collected in a receiver cooled in Dry-ice. For further details of the distillation apparatus see Smith [*J Am Chem Soc* **75** 184 1953]. **It is HIGHLY EXPLOSIVE; a strong protective screen should be used at all times.** [Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 318-320 1963.]

**Perrehenic acid (HRe<sub>2</sub>O<sub>4</sub>)** [13768-11-1] M 251.2,  $d^{25}$  2.16,  $pK^{25}$  -1.25. The acid is commercially available as a 65-70% w/w solution in H<sub>2</sub>O. It can be prepared freshly from a solution of pure finely powdered KRe<sub>2</sub>O<sub>4</sub> (10.0g, 35.6mmol) in hot H<sub>2</sub>O (120ml) at 100° which is placed on the top of a Dowex 50W-X1 cation exchange

resin column (40 x 20mm, 50g of 50-100 mesh with a 1cm layer of Pyrex wool). This resin is first washed with 6 M HCl (20ml) and flushed with pure hot H<sub>2</sub>O at 100° until the effluent is colourless and does not become turbid on addition of AgNO<sub>3</sub> solution (i.e. chloride free). The hot KRe<sub>2</sub>O<sub>4</sub> solution is applied to the resin in four 30ml portions (to keep the temperature high, boiling if necessary for the salt to remain in solution), and the hot column is rinsed with pure hot H<sub>2</sub>O. The rate of flow should be kept at 15-20ml/minute, the first effluent (*ca* 50ml) is discarded and the acidic fraction (to litmus, *ca* 150ml) is collected. This solution is concentrated to as small a volume as possible, preferably *in vacuo* over P<sub>2</sub>O<sub>5</sub>, and an aliquot is diluted and titrated against standard alkali to obtain the concentration of HRe<sub>2</sub>O<sub>4</sub> in the residue. It is a strong acid. It is then diluted with pure H<sub>2</sub>O as required. The theoretical amount of acid in the residue should be 8.68g. [Watt & Thompson *Inorg Synth* VII 187 1963.] It is also prepared by dissolving rhenium heptoxide (Re<sub>2</sub>O<sub>7</sub>, 1314-68-7) in H<sub>2</sub>O. [Smith & Long *J Am Chem Soc* 70 354 1948, Melaven et al. *Inorg Synth* III 188 1950.]

**Phosgene** [75-44-5] **M 98.9, m -118°, b 8.2°/756mm.** Dry the gas with Linde 4A molecular sieves, de-gas it and distil it under vacuum at low temperature. This should be done in a closed system such as a vacuum line. It is hydrolysed slowly by H<sub>2</sub>O, but does not fume in moist air. It is available in cylinders and as a ~20% solution in toluene. **It is HIGHLY TOXIC and should not be inhaled. If it is inhaled, the operator should lie still and, be made to breathe in ammonia vapour which reacts readily with phosgene to give urea.** [Pope et al. *J Chem Soc* 117 1410 1920, Beilstein 3 IV 41.]

**Phosphine** [7803-51-2] **M 34.0, m -133°, b -87.7°, pK<sup>25</sup> -14, pK<sub>b</sub> 28.** PH<sub>3</sub> is best purified in a gas line (in a vacuum) in an efficient fume cupboard. It is spontaneously flammable, has a strong odour of decayed fish and is **POISONOUS**. The gas is distilled through solid KOH towers (two), through a Dry ice-acetone trap (-78°, to remove H<sub>2</sub>O, and P<sub>2</sub>H<sub>4</sub> which spontaneously ignites with O<sub>2</sub>), then through two liquid N<sub>2</sub> traps (-196°), followed by distillation into a -126° trap (Dry-ice/methylcyclohexane slush), allowed to warm in the gas line and then sealed in ampoules preferably under N<sub>2</sub>. Its IR as  $\nu_{\max}$  at 327 (m), 1121 (m) and 900 (m) cm<sup>-1</sup>. [Klement in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 525-530 1963, Gokhale & Jolly *Inorg Synth* IX 56 1967.] PH<sub>3</sub> has also been absorbed into a solution of cuprous chloride in hydrochloric acid (when CuCl.PH<sub>3</sub> is formed). PH<sub>3</sub> gas is released when this solution is heated, and the gas is purified by passage through KOH pellets and then over P<sub>2</sub>O<sub>5</sub>. It ignites spontaneously in air with a luminous flame. Its solubility is 0.26ml/1 ml of H<sub>2</sub>O at 20°, and a crystalline *hydrate* is formed on releasing the pressure on an aqueous solution.

**Phosphonitrilic chloride (tetramer)** [1832-07-1] **M 463.9.** Purify it by zone melting, then recrystallise it from petroleum ether (b 40-60°) or *n*-hexane. [van der Huizen et al. *J Chem Soc, Dalton Trans* 1317 1986.]

**Phosphonitrilic chloride (trimer) (hexachlorocyclotriphosphazine)** [940-71-6] **M 347.7, m 112.8°, 113-114°.** Purify it by zone melting, by crystallisation from petroleum ether, *n*-hexane or \*benzene, and by sublimation. [van der Huizen et al. *J Chem Soc, Dalton Trans* 1311 1986, Meirovitch et al. *J Phys Chem* 88 1522 1984, Alcock et al. *J Am Chem Soc* 106 5561 1984; Winter & van de Grampel *J Chem Soc, Dalton Trans* 1269 1986.]

**Phosphonous acid (phosphonic acid, H-P=O(OH)<sub>2</sub>)** [13598-36-2] **M 82.0, m 73.6°, d<sub>4</sub><sup>25</sup> 1.651, pK<sub>1</sub> 1.43, pK<sub>2</sub> 6.67.** Phosphonous acid is best prepared by adding PCl<sub>3</sub> to concentrated HCl, when HCl gas is liberated and the solution is evaporated (fume-cupboard) until the temperature reaches 180° (all HCl gas is driven off). On cooling the acid separates as white hygroscopic (and deliquescent) crystals. It has a garlic taste, melts at 73.6°, and decomposes at 200° to give phosphine and phosphoric acid (note that PH<sub>3</sub> ignites in air with bright flashes, see above). It is a dibasic acid that slowly oxidises to phosphoric acid in air. Store it under dry N<sub>2</sub>, or as a 20% aqueous solution under N<sub>2</sub>. [Simon & Fehér *Z Anorg Allgem Chem* 230 298 1937, Voight & Gallais *Inorg Synth* IV 55 1953.]

**Tautomerism of phosphonous acid.** This acid is tautomeric with trihydroxyphosphine with the latter tautomer being in extremely small concentrations, e.g. 1 in 10<sup>12</sup>. The diester [e.g. (MeO)<sub>2</sub>P(H)=O] has the H linked to the P atom. The tautomerism is not like that of keto-enol tautomerism in carbon chemistry, and dialkylphosphites (RO)<sub>2</sub>P(H)=O do *not* react with sulfur to form the corresponding P-SH derivatives, but they do form solid salts such as (RO)<sub>2</sub>PONa which *do* add sulfur readily. Also this acid does form triesters such as (MeO)<sub>3</sub>P which is

trimethylphosphite or trimethoxyphosphine, and behave more like phosphines. [cf. F.A. Cotton, G. Wilkinson, C.A. Murillo and M. Bochmann, *Advanced Inorganic Chemistry*, 6<sup>th</sup> Edn, Interscience Publ, 1999, ISBN 0-471-19957-5.]

**Phosphoric acid** [7664-38-2] **M 98.0, m 42.3°, pK<sub>1</sub><sup>25</sup> 2.15, pK<sub>2</sub><sup>25</sup> 7.21, pK<sub>3</sub><sup>25</sup> 12.37.** Pyrophosphate can be removed from phosphoric acid by diluting with distilled H<sub>2</sub>O and refluxing overnight. By cooling to 11° and seeding with crystals obtained by cooling a few millilitres in a Dry-ice/acetone bath, 85% orthophosphoric acid crystallises as H<sub>3</sub>PO<sub>4</sub>·H<sub>2</sub>O. The crystals are collected on a sintered glass filter. [Weber & King *Inorg Synth I* 101 1939.]

**Phosphorus (red)** [7723-14-0] **M 31.0, m 590°/43atm, ignites at 200°, d<sub>4</sub><sup>20</sup> 2.34.** Heat it for 15 minutes in boiling distilled H<sub>2</sub>O, allow it to settle and wash it several times with boiling H<sub>2</sub>O. Transfer it to a Büchner funnel, wash it with hot H<sub>2</sub>O until the washings are neutral, then dry it at 100° and store it in a desiccator.

**Phosphorus (white)** [7723-14-0] **M 31.0, m 44.1°, b 287°, d<sub>4</sub><sup>20</sup> 1.82.** Purify white phosphorus by melting it under dilute H<sub>2</sub>SO<sub>4</sub>—dichromate (possible **carcinogen**) mixture and allow to stand for several days in the dark at room temperature. It remains liquid, and the initial milky appearance due to insoluble, oxidisable material gradually disappears. The phosphorus can then be distilled under vacuum in the dark [Holmes *Trans Faraday Soc* 58 1916 1962]. It sublimes *in vacuo*. Other methods of purification include extraction with dry CS<sub>2</sub> followed by evaporation of the solvent, or washing with 6M HNO<sub>3</sub>, then H<sub>2</sub>O, and drying under vacuum. It ignites in air at ~50°, or by friction if dry. Store and cut it under H<sub>2</sub>O. **POISONOUS, use gloves.**

**Phosphorus oxychloride** [10025-87-3] **M 153.3, b 105.5°, d<sub>4</sub><sup>20</sup> 1.675, n<sub>D</sub><sup>20</sup> 1.461.** Distil the liquid under reduced pressure to separate it from the bulk of the HCl and the phosphoric acid (from hydrolysis); the middle fraction is re-distilled into ampoules containing a little purified mercury. These ampoules are sealed and stored in the dark for 4–6 weeks with occasional shaking to facilitate reaction of any free chloride with the mercury. The POCl<sub>3</sub> is then again fractionally distilled and stored in sealed ampoules in the dark until required [Herber *J Am Chem Soc* 82 792 1960]. Lewis and Sowerby [*J Chem Soc* 336 1957] refluxed their distilled POCl<sub>3</sub> with Na wire for 4 hours, then removed the Na and again distilled. *Use Na only with almost pure POCl<sub>3</sub> to avoid explosions.* **HARMFUL VAPOURS; work in an efficient fume cupboard.**

**Phosphorus pentabromide** [7789-69-7] **M 430.6, m <100°, b 106°(dec).** Dissolve it in pure nitrobenzene at 60°, filtering off any insoluble residue on to sintered glass funnel, then allow it to crystallise by cooling. Wash the collected solid with dry Et<sub>2</sub>O and remove excess ether in a current of dry N<sub>2</sub>. (All manipulations should be performed in a dry-box.) [Harris & Payne *J Chem Soc* 3732 1958]. It fumes in moist air because of hydrolysis. **HARMFUL VAPOURS** (wash burning eyes with aqueous NaHCO<sub>3</sub>).

**Phosphorus pentachloride** [10026-13-8] **M 208.2, m 179-180°(sublimes).** [All operations should be carried out in an efficient fume cupboard.] Sublime it at 160-170° in an atmosphere of chlorine. Excess chlorine is then displaced by dry N<sub>2</sub> gas. All subsequent manipulations should be performed in a dry-box [Downs & Johnson *J Am Chem Soc* 77 2098 1955]. It fumes in moist air and attacks the eyes and the mucous membranes of the nose. It should not be breathed in and has very **HARMFUL VAPOURS** (wash burning eyes with aqueous NaHCO<sub>3</sub>).

**Phosphorus pentasulfide** [1314-80-3] **M 444.5, m 277-283°, 290°, b 513-515.** Purify P<sub>2</sub>S<sub>5</sub> by extraction and crystallisation with CS<sub>2</sub>, using a Soxhlet extractor, and is heated in a CO<sub>2</sub> atmosphere at 150° to remove solvent. It liberates H<sub>2</sub>S in moist air. **HARMFUL VAPOURS.** [Klements in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 568 1963.]

**Phosphorus pentoxide** [1314-56-3] **M 141.9, m 562°, b 605°.** It has been sublimed at 250° under vacuum into glass ampoules. It fumes in moist air and reacts violently with water. It is an excellent drying agent for use in desiccators. **HARMFUL VAPOURS and attacks skin, use gloves.** [Manley *J Chem Soc* 121 331 1922, Klements in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 541 1963.]

**Phosphorus sesquisulfide**  $P_4S_3$  [1314-85-8] **M 220.1, m 172°**. Extract  $P_4S_3$  with  $CS_2$ , filter it and evaporate it to dryness. *Alternatively*, place it in  $H_2O$ , and pass steam through it for an hour. The  $H_2O$  is then removed, the solid is dried, and recrystallised from  $CS_2$  (with the usual  $CS_2$  precautions) [Rogers & Gross *J Am Chem Soc* **74** 5294 1952].

**Phosphorus sulfochloride (phosphorus thiochloride)** [3982-91-0] **M 169.4, m -35°, b 122-124°, 125°(corr),  $d_4^{30}$  1.64,  $n_D^{30}$  1.556**. Possible impurities are  $PCl_5$ ,  $H_3PO_4$ ,  $HCl$  and  $AlCl_3$ . Gently mix it with  $H_2O$  to avoid a heavy emulsion; the product decolourises immediately and settles to the bottom layer. It is soluble in  $*C_6H_6$  and  $CCl_4$ . [Duval *Inorg Synth* **IV** 73 1953.] **HARMFUL VAPOURS**.

**Phosphorus tribromide** [7789-60-8] **M 270.7, m -41.5°, b 168-170°/725mm, 171-173°/atm, 172.9°/760mm(corr),  $d_4^{30}$  2.852**. It is decomposed by moisture, it should be kept dry and is *corrosive*. Purify it by distillation through an efficient fractionating column [see Whitmore & Lux *J Am Chem Soc* **54** 3451] in a slow stream of dry  $N_2$ , i.e. under strictly dry conditions. [Gay & Maxson *Inorg Synth* **II** 147 1946, *Org Synth Col Vol II* 358 1943.] Dissolve it in  $CCl_4$ , dry it over  $CaCl_2$ , filter and distil it. Store it in sealed ampoules under  $N_2$  and keep it away from light. It is also commercially available as a 1.0M solution in  $CH_2Cl_2$  ( $d^{25}$  1.488). **HARMFUL VAPOURS**.

**Phosphorus trichloride** [7719-12-2] **M 137.3, b 76°,  $d_4^{20}$  1.575, n 1.515**. Heat it under reflux to expel dissolved  $HCl$ , then distil it. It has been further purified by vacuum fractionation several times through a -45° trap into a receiver at -78°. [Forbes *Inorg Synth* **II** 145 1946.] **HARMFUL VAPOURS**.

**Phosphorus triiodide** [13455-01-1] **M 411.7, m 61°**. It decomposes in moist air and must be kept in a desiccator over  $CaCl_2$ . It is crystallised from sulfur-free  $CS_2$ ; otherwise the **m** decreases to *ca* 55°. It is best to prepare it freshly. [Germann & Traxler *J Am Chem Soc* **49** 307 1927, Klement in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 541 1963.] **HARMFUL VAPOURS**.

**12-Phosphotungstic acid** [12501-23-4] **M 2880.2, m ~96°**. A few drops of conc  $HNO_3$  are added to 100g of phosphotungstic acid dissolved in 75ml of water, in a separating funnel, and the solution is extracted with diethyl ether. The lowest of the three layers, which contains a phosphotungstic acid-ether complex, is separated, washed several times with 2M  $HCl$ , then with water and again extracted with ether. Evaporation of the ether under vacuum with mild heating on a water bath gives crystals which are dried under vacuum and ground [Matijevic & Kerker, *J Am Chem Soc* **81** 1307 1959].

**Platinum (IV) ammonium chloride [ammonium hexachloroplatinate (IV),  $(NH_4)_2 PtCl_6$ ]** [16919-58-7] **M 443.9**. This salt is a useful Pt compound and is a very good source for preparing many amine complexes. It provides spongy Pt at high temperature. It is prepared by dissolving the less expensive spongy Pt (9.76g, 50mmol) in *aqua regia* (200ml,  $1HNO_3:4HCl$  by volume, and allowed to stand for 30 minutes until orange-red in colour) by heating and concentrating in an evaporating dish (*use an efficient fume hood and care as it is a dangerous operation*) to a thick syrup which is then taken almost to dryness on a water bath. Avoid drying at higher temperatures as it becomes incompletely soluble in  $H_2O$ . The orange-red crusty residue is dissolved in  $H_2O$  (100ml), filtered; and to it is add slowly a solution of  $NH_4Cl$  (8g, 150mmol) in  $H_2O$  (100ml) with stirring, when a yellow finely crystalline precipitate separates. Set the mixture aside for 30 minutes in an ice bath. Filter off the platinate salt and wash it several times with 1% aqueous  $NH_4Cl$  (30ml portions), then with  $EtOH$  (which removes any corresponding Pd salt impurity) and  $Et_2O$ . Dry the salt first in air, powder the crystals and dry further at 120° for 1 hour to provide an almost quantitative yield. Note that although the filtrates may be yellow in colour they only contain minute traces of Pt as the colour can be imparted at as low concentrations as 1 part in 20,000 parts of  $H_2O$ . The salt is not hygroscopic. Its solubility (by weight) in  $H_2O$  is 0.5% at 20° and 3.37% at 100°, and it is still less soluble in 1M  $NH_4Cl$  (0.0028%). It decomposes above 185° and yields *spongy Pt* at higher temperatures. It is reduced to the metal by  $H_2$  at 120° and yields  $H_2PtCl_6$  with  $Cl_2$ . [Kauffman *Inorg Synth* **9** 182 1967.]

**Platinum (II) chloride** [10025-65-7] **M 266.0, d<sub>4</sub><sup>20</sup> 5.87**. It is purified by heating at 450° in a stream of Cl<sub>2</sub> for 2 hours. Some sublimation occurs because the PtCl<sub>2</sub> sublimes completely at 560° as red (almost black) needles. This sublimate can be combined to the bulk chloride, and while still at *ca* 450° it should be transferred to a container and cooled in a desiccator. A probable impurity is PtCl<sub>4</sub>. To test for this add a few drops of H<sub>2</sub>O (in which PtCl<sub>4</sub> is soluble) to the salt, filter and add an equal volume of saturated aqueous NH<sub>4</sub>Cl to the filtrate. If no precipitate is formed within 1 minute, then the product is pure. If a precipitate appears, then the whole material should be washed with small volumes of H<sub>2</sub>O until the soluble PtCl<sub>4</sub> is removed. The purified PtCl<sub>2</sub> is partly dried by suction and then dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. It is insoluble in H<sub>2</sub>O, but soluble in HCl to form chloroplatinic acid (H<sub>2</sub>PtCl<sub>4</sub>) by disproportionation. [Cohen *Inorg Synth* VI 209 1960.]

**Potassium bicarbonate** [298-14-6] **M 100.1**. It is crystallised from water at 65-70° (1.25ml/g) by filtering and then cooling to 15° (~0.4ml/g). During all operations, CO<sub>2</sub> is passed through the stirred mixture. The crystals are sucked dry at the pump, washed with distilled water, dried in air and then over H<sub>2</sub>SO<sub>4</sub> in an atmosphere of CO<sub>2</sub>. It is much less soluble than the carbonate in H<sub>2</sub>O (see below).

**Potassium biiodate** [13455-24-8] **M 389.9**. Crystallise the biiodate three times from hot water (3ml/g), and stirring continuously during each cooling. After drying at 100° for several hours, the crystals are suitable for use in volumetric analysis.

**Potassium bisulfate** [7646-93-7] **M 136.2, m 214°, 218°**. Crystallise it from H<sub>2</sub>O (1ml/g) between 100° and 0°. It is also formed when a warm solution of K<sub>2</sub>SO<sub>4</sub> in conc H<sub>2</sub>SO<sub>4</sub> is cooled down.

**Potassium borohydride** [13762-51-1] **M 53.9, m ~500°(dec)**. Crystallise it from liquid ammonia. It is slowly hydrolysed by H<sub>2</sub>O. Its solubility at ~20° in H<sub>2</sub>O or liquid NH<sub>3</sub> is 20%, in MeOH it is 0.7%, in Me<sub>2</sub>NCHO it is 15% and in MeOH/H<sub>2</sub>O (1:4) it is 13%. [Jons & Wallbridge *Progr Inorg Chem* 11 99-231 1970.]

**Potassium bromate** [7758-01-2] **M 167.0, m 350°(dec at 370°), d<sub>4</sub><sup>20</sup> 3.27**. Crystallise KBrO<sub>3</sub> from distilled H<sub>2</sub>O (2ml/g) between 100° and 0°. To remove bromide contamination, a 5% solution in distilled H<sub>2</sub>O, cooled to 10°, is bubbled with gaseous chlorine for 2 hours, then filtered and extracted with reagent grade CCl<sub>4</sub> until colourless and odourless. After evaporating the aqueous phase to about half its volume, it is cooled again slowly to about 10°. The crystalline KBrO<sub>3</sub> that separates, is washed with 95% EtOH and dried in a vacuum [Boyd et al. *J Am Chem Soc* 74 237 1952]. Another way to remove Br<sup>-</sup> ions is by stirring several times in MeOH and then drying at 150° [Field & Boyd *J Phys Chem* 89 3767 1985].

**Potassium bromide** [7758-02-3] **M 119.0, m 734°, d<sub>4</sub><sup>20</sup> 2.75**. Crystallise the bromide from distilled water (1ml/g) between 100° and 0°. Wash it with 95% EtOH, followed by Et<sub>2</sub>O. Dry it in air, then heat it at 115° for 1 hour, pulverise it, then heat it in a vacuum oven at 130° for 4 hours. It has also been crystallised from aqueous 30% EtOH, or EtOH, and dried over P<sub>2</sub>O<sub>5</sub> under vacuum before heating in an oven.

**Potassium carbonate** [584-08-7] **M 138.2, m 898°, d<sub>4</sub><sup>20</sup> 2.3**. It crystallises from water between 100° and 0°. The solubility in H<sub>2</sub>O is 105% at 0°, 127% at 60° and 205% at 135° (**b** of saturated solution). After two recrystallisations of technical grade material, it had B, Li and Fe at 1.0, 0.04 and 0.01 ppm, respectively. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 987 1963.]

**Potassium chlorate** [3811-04-9] **M 122.6, m 368°**. It has been recrystallised from water (1.8ml/g) between 100° and 0°, and the crystals were filtered onto sintered glass. Keep away from organic material as it oxidises them readily.

**Potassium chloride** [7447-40-7] **M 74.6, m 771°, d<sub>4</sub><sup>20</sup> 1.98**. Dissolve it in conductivity water, filter it, and saturate it with chlorine (generated from conc HCl and KMnO<sub>4</sub>). Excess chlorine is boiled off, and the KCl is precipitated by HCl (generated by dropping conc HCl into conc H<sub>2</sub>SO<sub>4</sub>). The precipitate is washed with water, dissolved in conductivity water at 90-95°, and crystallised by cooling to about -5°. The crystals are drained at the centrifuge, dried in a vacuum desiccator at room temperature, then fused in a platinum dish under N<sub>2</sub>, cooled

and stored in a desiccator. Potassium chloride has also been sublimed in a stream of pre-purified  $N_2$  gas and collected by electrostatic discharge [Craig & McIntosh *Can J Chem* **30** 448 1952].

**Potassium chromate** [7789-00-6] **M 194.2, m 975°**,  $d_4^{20}$  2.72,  $pK_1^{25}$  0.74,  $pK_2^{25}$  6.49 (for  $H_2CrO_4$ ). Crystallise it from conductivity water (0.6g/ml at 20°), and dry it between 135° and 170°.

**Potassium cobalticyanide** [13963-58-1] **M 332.4, m dec on heating**,  $d_4^{25}$  1.878. Crystallise it from water to remove traces of HCN. Its solubility in 87-88% EtOH is 1 in 7500 at 20°. [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1542 1965.]

**Potassium cyanate** [590-28-3] **M 81.1,  $d_4^{20}$  2.05,  $pK^{25}$  3.46 (for HCNO)**. Common impurities include ammonia and bicarbonate ion (from hydrolysis). Purify it by preparing a saturated aqueous solution at 50°, neutralising with acetic acid, filtering, adding two volumes of EtOH and keeping for 3-4 hours in an ice bath. (More EtOH can lead to co-precipitation of  $KHCO_3$ .) Filter, wash it with EtOH and dry it rapidly in a vacuum desiccator ( $P_2O_5$ ). The process is repeated [Vanderzee & Meyers *J Chem Soc* **65** 153 1961].

**Potassium cyanide** [151-50-8] **M 65.1, m 634°**,  $d_4^{20}$  1.52. A saturated solution in  $H_2O$ -ethanol (1:3) at 60° is filtered and cooled to room temperature. Absolute EtOH is added, with stirring, until crystallisation ceases. The solution is again allowed to cool to room temperature (during 2-3 hours), then the crystals are filtered off, washed with absolute EtOH, and dried, first at 70-80° for 2-3 hours, then at 105° for 2 hours [Brown et al. *J Phys Chem* **66** 2426 1962]. It has also been purified by melting in a vacuum and by zone refining. Its solubility is 33% ( $H_2O$ ), 66% (boiling  $H_2O$ ), 33% (glycerol), 4% (MeOH) and 1% (EtOH). It hydrolyses in  $H_2O$ , and a 0.1M aqueous solution of KCN has a pH of 11, i.e.  $CO_2$  in the  $H_2O$  makes it acidic enough to liberate a lot of the more poisonous HCN. To keep the HCN in solution it must be made quite alkaline (pH <11).  $LD_{50}$  oral toxicity in rat is ~10mg/Kg. **HIGHLY POISONOUS work in an efficient fume cupboard.**

**Potassium dichromate** [7778-50-9] **M 294.2, m 398°(dec)**,  $d_4^{20}$  2.68. Crystallise it from water (g/ml) between 100° and 0° and dry it under vacuum at 156°. (Possible **CARCINOGEN**.)

**Potassium dihydrogen phosphate** [7778-77-0] **M 136.1**. Dissolve it in boiling distilled water (2ml/g), keep on a boiling water-bath for several hours, then filter it through paper pulp to remove any turbidity. Cool rapidly with constant stirring, and the crystals are collected on to hardened filter paper, using suction, washed twice with ice-cold water, once with 50% EtOH, and dried at 105°. Alternative crystallisations are from water, then 50% EtOH, and again water, or from concentrated aqueous solution by addition of EtOH. It is freed from traces of Cu by extracting its aqueous solution with diphenylthiocarbazon in  $CCl_4$ , followed by repeated extraction with  $CCl_4$  to remove traces of diphenylthiocarbazon.

**Potassium dithionate** [13455-20-4] **M 238.3,  $pK_{Est(1)}$  -3.4,  $pK_2^{25}$  0.49 (for dithionic acid)**. Crystallise it from water (1.5ml/g) between 100° and 0°.

**Potassium ferricyanide** [13746-66-2] **M 329.3,  $pK^{25}$  <1 (for ferricyanide)**. It has been recrystallised repeatedly from hot water (1.3ml/g) and dried under vacuum in a desiccator.

**Potassium ferrocyanide trihydrate** [14459-95-1] **M 422.4,  $pK_3^{25}$  2.57,  $pK_4^{25}$  4.35 (for ferrocyanide)**. It is purified by repeated crystallisation from distilled water, and never heating above 60°. The *anhydrous* salt is prepared by drying at 110° over  $P_2O_5$  in a vacuum desiccator. To obtain the *trihydrate*, it is necessary to equilibrate the salt in a desiccator over a saturated aqueous solution of sucrose and NaCl. It can also be precipitated from a saturated solution at 0° by adding an equal volume of cold 95% EtOH, setting aside for several hours, then centrifuge and wash with cold 95% EtOH. It is finally sucked air dry with water-pump vacuum. The *anhydrous* salt is obtained by drying the hydrate in a platinum boat at 90° in a slow stream of  $N_2$  [Loftfield & Swift *J Am Chem Soc* **60** 3083 1938].

**Potassium fluorosilicate** [16871-90-2] **M 220.3,  $d_4^{20}$  2.3,  $pK^{25}$  1.92 (for  $H_2SiF_6$ )**. Crystallise it several times from conductivity water (100ml/g) between 100° and 0°.

**Potassium hexachloroiridate (III) ( $K_3IrCl_6$ )** [14024-41-1] **M 483.1**. Crystallise it from hot aqueous solution, and the solution should be olive-green in colour. If it has a tinge of red, then some would have oxidised to ( $K_2IrCl_6$ ) [see following entry]. In this case make a concentrated solution in  $H_2O$ , and bubble  $H_2S$  through until the solution is clearly olive-green in colour due to Ir(III). Add KCl and  $K_3IrCl_6 \cdot 3H_2O$  deposits on evaporating under  $N_2$ . Filter it off, wash it with a little  $H_2O$ , then EtOH and dry it *in vacuo* away from air. [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II pp 1595 1965.]

**Potassium hexachloroiridate (IV) ( $K_2IrCl_6$ )** [16920-56-2] **M 522.2**. Crystallise it from hot aqueous solution containing a few drops of  $HNO_3$  to keep it in the oxidised state. It forms small shiny red-black octahedral crystals which are dried at  $100^\circ$  and give a red powder on grinding [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II pp 1593-1594 1965].

**Potassium hexachloroosmate (IV)** [16871-60-6] **M 481.1**. Crystallise it from hot dilute aqueous HCl. [Turner et al. *Anal Chem* **30** 1708 1958.]

**Potassium hexachloroplatinate (IV)** [16921-30-5] **M 486.0, m 250<sup>o</sup>(dec)**. It crystallises from water (20ml/g) between 100 and  $0^\circ$ . Its solubility in  $H_2O$  is 0.7% at  $0^\circ$ , 1.12% at  $20^\circ$ , 2.16% at  $50^\circ$  and 5.13% at  $100^\circ$ . [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1571 1965.]

**Potassium hexacyanochromate (III) ( $K_3[Cr(CN)_6] \cdot 3H_2O$ )** [13601-11-1] **M 418.5**. It forms yellow crystals from water. Recrystallise it two or three times from  $H_2O$  and dry it over  $H_2SO_4$ . Its solubility at  $20^\circ$  is 30.96g/100g  $H_2O$ , and it is insoluble in EtOH. Aqueous solutions tend to decompose especially in the presence of light or on heating when  $Cr(OH)_3$  separates. [Hein & Herzog in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II pp 1373-1374 1965.]

**Potassium hexafluorophosphate** [17084-13-8] **M 184.1,  $pK_1^{25} \sim 0.5$ ,  $pK_2^{25} 5.12$  (for fluoro-phosphoric acid  $H_2PO_3F$ )**. Crystallise it from alkaline aqueous solution, using polyethylene vessels, or from 95% EtOH, and dry it in a vacuum desiccator over KOH. [Kloditz *Z Anorg Allgem Chem* **284** 144 1956, Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 196 1963.]

**Potassium hexafluorozirconate ( $K_2ZrF_6$ )** [16923-95-8] **M 283.4,  $d_4^{20} 3.48$** . Recrystallise it from hot water (solubility is 0.78% at  $2^\circ$  and 25% at  $100^\circ$ ).

**Potassium hydrogen fluoride** [7789-29-9] **M 78.1, m 225<sup>o</sup>(dec)**. It crystallises from water. It is very soluble in hot  $H_2O$  and 41% at  $21^\circ$ . [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 237 1963.]

**Potassium hydroxide (solution)** [1310-58-3] **M 56.1,  $pK^{25} 16$  (for aquo  $K^+$ )**. Its carbonate content can be reduced by rinsing KOH sticks rapidly with water prior to dissolving them in boiled out distilled water. Alternatively, a slight excess of saturated  $BaCl_2$  or  $Ba(OH)_2$  can be added to the solution which, after shaking well, is set aside so that the  $BaCO_3$  is allowed to separate out. Davies and Nancollas [*Nature* **165** 237 1950] rendered KOH solutions carbonate free by ion exchange using a column of Amberlite IR-100 in the  $OH^-$  form.

**Potassium iodate** [7758-05-6] **M 214.0,  $pK^{25} 0.80$  (for  $HIO_3$ )**. It has been crystallised twice from distilled water (3ml/g) between 100 and  $0^\circ$ , dried for 2 hours at  $140^\circ$  and cooled in a desiccator. Analytical reagent grade material dried in this way is suitable for use as an analytical standard.

**Potassium iodide** [7681-11-0] **M 166.0,  $pK^{25} -8.56$  (for HI)**. Crystallise it from distilled water (0.5ml/g) by filtering the near-boiling solution and cooling. To minimise oxidation to iodine, the process can be carried out under  $N_2$  and the salt is dried under a vacuum over  $P_2O_5$  at  $70-100^\circ$ . Before drying, the crystals can be washed with EtOH or with acetone followed by petroleum ether. It has also been recrystallised from water/ethanol. After 2 recrystallisations, ACS/USP grade had Li and Sb at  $<0.02$  and  $<0.01$  ppm respectively. [Lingane & Kolthoff *Inorg Synth* **I** 163 1939.]

**Potassium nickel sulfate hexahydrate** [13842-46-1] **M 437.1**. Crystallise it from H<sub>2</sub>O (1.7ml/g) between 75 and 0°.

**Potassium nitrate (saltpetre)** [7757-79-1] **M 101.1, m 336°**. It crystallises from hot H<sub>2</sub>O (0.5ml/g) on cooling (cf KNO<sub>2</sub> below). Dry it for 12 hours under vacuum at 70°. The solubility in H<sub>2</sub>O is 13.3% at 0°, 110% at 60°, and 246% at 100°. After two recrystallisations, technical grade salt had <0.001 ppm of metals. The fused salt is a powerful oxidising agent.

**Potassium nitrite** [7758-09-0] **M 85.1, m 350°(dec), pK<sup>20</sup> 3.20 (for HNO<sub>2</sub>)**. A saturated solution at 0° is warmed and partially evaporated under vacuum. The crystals so obtained are filtered off from the warm solution. (This procedure is designed to reduce the level of nitrate impurity and is based on the effects of temperature on solubility. The solubility of KNO<sub>3</sub> in water is 13g/100ml at 0°, 247g/100ml at 100°; for KNO<sub>2</sub> the corresponding figures are 280g/100ml and 413g/100ml.) *Alternatively*, dissolve it in H<sub>2</sub>O and precipitate by adding of EtOH.

**Potassium nitrosodisulfonate (Fremy's Salt)** [14293-70-0] **M 268.3**. It forms yellow needles (dimeric) which dissolve in H<sub>2</sub>O to give the violet monomeric free radical. It is purified by dissolving (~12g) in 2M KOH (600ml) at 45°, filtering the blue solution and keeping it in a refrigerator overnight. The golden yellow crystals (10g) are filtered off, washed with MeOH (3x), then Et<sub>2</sub>O and stored in a glass container in a vacuum over KOH. It is stable indefinitely when dry. [Cram & Reeves *J Org Chem* **80** 3094 1958, Schenk *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 505 1963.]

**Potassium osmate (VI) dihydrate** [19718-36-6] **M 368.4**. It forms *hygroscopic POISONOUS* crystals which are soluble in H<sub>2</sub>O but insoluble in EtOH and Et<sub>2</sub>O. It decomposes slowly in H<sub>2</sub>O to form the *tetroxide* which attacks the eyes. The solid should be kept dry and in this form it is relatively safe. [Lloyd et al. *Synthesis* 610 1972, Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1604 1965.]

**Potassium perchlorate** [7778-74-7] **M 138.6, m 400(dec), d<sub>4</sub><sup>20</sup> 2.52, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. It crystallises from boiling water (5ml/g) on cooling. Dry it under vacuum at 105°.

**Potassium periodate (potassium metaperiodate)** [7790-21-8] **M 230.0, m 582°, d<sub>4</sub><sup>20</sup> 3.62**. Crystallise it from distilled water. Its solubility in H<sub>2</sub>O is 0.2% at 0°, 0.4% at 20°, 4.4% at 80° and 7.9% at 100°. [Hill *Inorg Synth* I 171 1939, Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 325 1963.]

**Potassium permanganate** [7722-64-7] **M 158.0, m 240(dec), d<sub>4</sub><sup>20</sup> 2.7, pK<sup>25</sup> -2.25 (for HMnO<sub>4</sub>)**. Crystallise it from hot water (19ml/g at 15°, 4ml/g at 65°), then dry it in a vacuum desiccator over CaSO<sub>4</sub>. Phillips and Taylor [*J Chem Soc* 4242 1962] cooled an aqueous solution of KMnO<sub>4</sub>, saturated at 60°, to room temperature in the dark, and filtered it through a No.4 porosity sintered-glass filter funnel. The solution was allowed to evaporate in air in the dark for 12 hours, and the supernatant liquid was decanted from the crystals, which were dried as quickly as possible with filter paper. It is a secondary analytical standard. It is a useful oxidant and is used as a topical antiseptic.

**Potassium peroxydisulfate (potassium persulfate, K<sub>2</sub>S<sub>2</sub>O<sub>8</sub>)** [7727-21-1] **M 270.3**. Crystallise the persulfate twice from distilled water (10ml/g) and dry it at 50° in a vacuum desiccator. Its solubility in H<sub>2</sub>O is 1.6% at 0°, 4.5% at 20°, and 7.2% at 30°. An aqueous solution decomposes on long standing with evolution of O<sub>2</sub> and formation of KHSO<sub>4</sub>. It is a powerful oxidising agent. Store it at ~10°. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 390 1963.]

**Potassium peroxymonosulfate (Oxone, potassium monopersulfate triple salt; 2KHSO<sub>5</sub>·KHSO<sub>4</sub>·K<sub>2</sub>SO<sub>4</sub>)**, [37222-66-5, 70693-62-8 (*triple salt*)] **M 614.8**. This is a stable form of Caro's acid and should contain >4.7% of active oxygen. It can be used in EtOH/H<sub>2</sub>O and EtOH/AcOH/H<sub>2</sub>O solutions. If active oxygen is too low, it is

best to prepare it afresh from 1mole of  $\text{KHSO}_5$ , 0.5mole of  $\text{KHSO}_4$  and 0.5mole of  $\text{K}_2\text{SO}_4$ . [Kennedy & Stock *J Org Chem* **25** 1901 1960, Stephenson US Patent 2,802,722 1957.] A rapid preparation of *Caro's acid* is made by stirring finely powdered potassium persulfate (M 270.3) into ice-cold conc  $\text{H}_2\text{SO}_4$  (7ml) and when homogeneous add ice (40-50g). It is stable for several days if kept cold. Keep away from organic matter as it is a **STRONG OXIDANT**. A detailed preparation of *Caro's acid* (*hypersulfuric acid*,  $\text{H}_2\text{SO}_5$ , [7722-86-3]) in crystalline form m  $-45^\circ$  from  $\text{H}_2\text{O}_2$  and chlorosulfonic acid was described by Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 388 1963.

**Potassium perrhenate ( $\text{KReO}_4$ )** [10466-65-6] M 289.3, m  $550^\circ$ , b 1360-1370 $^\circ$ , d<sub>4</sub><sup>25</sup> 4.887, n<sub>D</sub><sup>25</sup> 1.643. pK<sup>25</sup> -1.25 (for  $\text{HReO}_4$ ). It forms white tetrahedral crystals from  $\text{H}_2\text{O}$  where its solubility is 12.1g/L at  $20^\circ$  and 140g/L at  $100^\circ$ . The salt is then fused in a platinum crucible in air at  $750^\circ$ . [Smith & Long *J Am Chem Soc* **70** 354 1948, Watt & Thompson *Inorg Synth* **7** 187 1963.]

**Potassium reineckate** [34430-73-4] M 357.5. Crystallise it from  $\text{KNO}_3$  solution, then from warm water [Adamson *J Am Chem Soc* **80** 3183 1958].

**Potassium(VI) ruthenate** [31111-21-4] M 243.3. Dissolve the ruthenate in  $\text{H}_2\text{O}$  and evaporate until crystals are formed. The crystals are iridescent green prisms which appear red in thin films. A possible impurity is  $\text{RuO}_4$ ; in this case wash with  $\text{CCl}_4$  (which dissolves  $\text{RuO}_4$ ). The concentration of an aqueous solution of  $\text{RuO}_4^{2-}$  (orange colour) can be estimated from the absorbance at 385nm ( $\epsilon$  1030  $\text{M}^{-1} \text{cm}^{-1}$ ), or at 460nm ( $\epsilon$  1820  $\text{M}^{-1} \text{cm}^{-1}$ ). [Lee et al. *Can J Chem* **50** 3741 1972, Connic & Hurley *J Am Chem Soc* **74** 5012 1952, Grube et al. *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1600 1965].

**Potassium selenocyanate** [3425-46-5] M 144.1. Dissolve it in acetone, filter and precipitate it by adding  $\text{Et}_2\text{O}$ .

**Potassium sulfate** [7778-80-5] M 174.3, m  $1069^\circ$ , d<sub>4</sub><sup>20</sup> 2.67. It crystallised from distilled water (4ml/g at  $20^\circ$ ; 8ml/g at  $100^\circ$ ) between  $100^\circ$  and  $0^\circ$ .

**Potassium tetrachloroplatinate(II)** [10025-99-7] M 415.1, m  $500^\circ(\text{dec})$ . It forms crystals from aqueous 0.75M  $\text{HCl}$  (20ml/g) between  $100^\circ$  and  $0^\circ$ . Wash them with ice-cold water and dry.

**Potassium tetracyanopalladate (II) trihydrate** [10025-98-6] M 377.4. All operations should be carried out in an efficient fume cupboard. **Cyanide is very POISONOUS**. Dissolve the complex (ca 5g) in a solution of  $\text{KCN}$  (4g) in  $\text{H}_2\text{O}$  (75ml) with warming and stirring, and evaporate hot till crystals appear. Cool, filter off the crystals and wash them with a few drops of cold  $\text{H}_2\text{O}$ . Further concentration of the mother liquors provides more crystals. The complex is recrystallised from  $\text{H}_2\text{O}$  as the colourless *trihydrate*. It *effloresces* in dry air and dehydrates at  $100^\circ$  to the *monohydrate*. The *anhydrous salt* is obtained by heating at  $200^\circ$ , but at higher temperatures it decomposes to  $(\text{CN})_2$ , Pd and  $\text{KCN}$ . [Bigelow *Inorg Synth* **II** 245 1946.]

**Potassium tetrafluoroborate (potassium borofluoride)** [14075-53-7] M 125.9, m  $530^\circ$ , d<sub>4</sub><sup>30</sup> 2.505, pK<sup>25</sup> -4.9 (for  $\text{HBF}_4$ ). Crystallise it from  $\text{H}_2\text{O}$  (solubility% (temperature): 0.3 ( $3^\circ$ ), 0.45 ( $20^\circ$ ), 1.4 ( $40^\circ$ ), 6.27 ( $100^\circ$ ), and dry it under vacuum. It is a non-hygroscopic salt. A 10% solution is transparent blue at  $100^\circ$ , green at  $90^\circ$  and yellow at  $60^\circ$ . [Vörlander et al. *Chem Ber* **65** 535 1932, Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 223 1963.]

**Potassium thiocyanate** [333-20-0] M 97.2, m  $172^\circ$ , pK<sup>25</sup> -1.85 (for  $\text{HSCN}$ ). Crystallise it from  $\text{H}_2\text{O}$  if much chloride ion is present in the salt, otherwise from  $\text{EtOH}$  or  $\text{MeOH}$  (optionally by addition of  $\text{Et}_2\text{O}$ ). Filter off on a Büchner funnel without paper, and dry it in a desiccator at room temperature before heating for 1 hour at  $150^\circ$ , with a final 10-20 minutes at  $200^\circ$  to remove the last traces of solvent [Kolthoff & Lingane *J Am Chem Soc* **57** 126 1935]. Store it in the dark.

**Potassium thiosulfate hydrate** [13446-67-8, 10294-66-3 (75% aqueous solution)] **M 190.3**,  $\text{pK}_1^{25}$  **0.6**,  $\text{pK}_2^{25}$  **1.74** (for  $\text{H}_2\text{S}_2\text{O}_3$ ). Crystallise it from warm water (0.5ml/g) by cooling in an ice-salt mixture. It is a good reducing agent used in analytical chemistry. [Foerster & Mommen *Chem Ber* **57** 258 1924.]

**Potassium tungstate (ortho dihydrate)** [37349-36-3; 7790-60-5] **M 362.1**, **m 921°**,  $d_4^{20}$  **3.12**,  $\text{pK}_1^{25}$  **2.20**,  $\text{pK}_2^{25}$  **3.70** (for  $\text{H}_2\text{WO}_4$ ). Crystallise it from hot water (0.7ml/g).

**Praseodymium trichloride hexahydrate** [10361-79-2] **M 355.4**,  $\text{pK}_1^{25}$  **8.55** (for  $\text{Pr}^{3+}$  hydrol). Its 1M solution in 6M HCl is passed twice through a Dowex-1 anion-exchange column. The eluate is evaporated in a vacuum desiccator to half its volume and allowed to crystallise [Katzin & Gulyas *J Phys Chem* **66** 494 1962].

**Praseodymium oxide ( $\text{Pr}_6\text{O}_{11}$ )** [12037-29-5] **M 1021.4**. Dissolve the oxide in acid (perchloric acid), precipitate it as the oxalate and the salt is ignited at 650° to give the oxide.

**Reinecke salt.** See ammonium reineckate above.

**Rhodium (I) carbonyl chloride (di- $\mu$ -chloro-tetracarbonyl dirhodium I,  $[\text{Rh}(\text{CO})_2\text{Cl}]_2$ )** [14532-22-9] **M 388.8**, **m 121°**, **124-125°**. This catalyst is soluble in most organic solvents, but not petroleum ethers, and forms orange-red crystals from hexane. It sublimes at 80°/0.1mm to a red solid. It decomposes on exposure to air when in organic solvents but the solid is stable in dry air. It is moisture sensitive and should be stored in ampoules under  $\text{N}_2$  or argon. It catalyses ring-opening silylformylation of olefins. [McCleverty & Wilkinson *Inorg Synth* **8** 211 1966, Dahl et al. *J Am Chem Soc* **83** 1761 1961, Cramer *Inorg Synth* **15** 14 1974, Fukumoto et al. *J Org Chem* **58** 4187 1993, Colton et al. *Aust J Chem* **23** 1351 1970.]

**Rhodium (III) chloride** [10049-07-7] **M 209.3**, **m >100°(dec)**, **b 717°**. Probable impurities are KCl and HCl. Wash the chloride well with small volumes of  $\text{H}_2\text{O}$  to remove excess KCl and KOH and dissolve it in the minimum volume of conc HCl. Evaporate it to dryness on a steam bath to give wine-red coloured  $\text{RhCl}_3 \cdot 3\text{H}_2\text{O}$ . Leave it on the steam bath until the odour of HCl is lost – do not try to dry further as it begins to decompose above 100° to the oxide and HCl. It is not soluble in  $\text{H}_2\text{O}$  but soluble in alkalis or CN solutions and forms double salts with alkali chlorides. [Anderson & Basolo *Inorg Synth* **VII** 214 1963.]

**Rubidium bromide** [7789-39-1] **M 165.4**, **m 682°**, **b 1340°**,  $d_4^{20}$  **3.35**. The bromide is a white crystalline powder which crystallises from  $\text{H}_2\text{O}$  (solubility: 50% in cold and 67% in boiling  $\text{H}_2\text{O}$  to give a neutral solution). It also crystallises from near-boiling water (0.5ml/g) by cooling to 0°.

**Rubidium chlorate** [13446-71-4] **M 168.9**,  $d_4^{20}$  **3.19**. It crystallises from water (1.6ml/g) by cooling from 100°.

**Rubidium chloride** [7791-11-9] **M 120.9**, **m 715°**, **m 1383°**,  $d_4^{20}$  **2.80**. Crystallise it from water (0.7ml/g) by cooling to 0° from 100°. Its solubility in  $\text{H}_2\text{O}$  is 77.3% at 0.6°, 90.3% at 70° and 147% at boiling point. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 951-955 1963.]

**Rubidium nitrate** [13126-12-0] **M 147.5**, **m 305°**,  $d_4^{20}$  **3.11**. Crystallise the nitrate from hot water (0.25ml/g) by cooling to room temperature.

**Rubidium perchlorate** [13510-42-4] **M 184.9**,  $d_4^{20}$  **2.80**,  $\text{pK}^{25}$  **-2.4 to -3.1** (for  $\text{HClO}_4$ ). Crystallise the perchlorate from hot water (1.6ml/g) by cooling to 0°.

**Rubidium sulfate** [7488-54-2] **M 267.0**, **m 1050°**,  $d_4^{20}$  **6.31**. Crystallise the sulfate from water (1.2ml/g) between 100° and 0°.

**Ruthenium (III) chloride dihydrate ( $\beta$ -form)** [14898-67-0] **M 207.4 +  $\text{H}_2\text{O}$** , **m >500°(dec)**,  $d_4^{20}$  **3.11**,  $\text{pK}_1^{25}$  **3.40** (for aquo  $\text{Rh}^{3+}$  hydrolysis). Dissolve the salt in  $\text{H}_2\text{O}$ , filter and concentrate to crystallisation in the

absence of air to avoid oxidation. Evaporate the solution in a stream of HCl gas while being heated just below its boiling point until a syrup is formed and finally to dryness at 80-100° and dried in a vacuum over H<sub>2</sub>SO<sub>4</sub>. When heated at 700° in the presence of Cl<sub>2</sub> the insoluble  $\alpha$ -form is obtained [Grube in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1598 1965, Carlsen et al. *J Org Chem* **46** 3936 1981].

**Ruthenium (IV) oxide** [12036-10-1] M 133.1, d<sub>4</sub><sup>20</sup> 6.97. Free the oxide from nitrates by boiling in distilled water and filtering. A more complete purification is based on fusion in a KOH/KNO<sub>3</sub> mix to form the soluble ruthenate and perruthenate salts. The melt is dissolved in water, and filtered, then acetone is added to reduce the ruthenates to the insoluble hydrated oxide which, after making a slurry with paper pulp, is filtered and ignited in air to form the *anhydrous* oxide [Campbell et al. *Anal Chem* **33** 58 1961].

**Samarium (II) iodide (SmI<sub>2</sub>)** [32248-43-4] M 404.2, m 520°, b 1580. A possible impurity is SmI<sub>3</sub> from which it is made. If present, grind the solid to a powder and heat it in a stream of pure H<sub>2</sub>. The temperature (~ 500-600°) should be below the m (~628°) of SmI<sub>3</sub>, since the molten compounds react very slowly. [Wetzel in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II pp 1149, 1150 1965.] It promotes a pseudo-Barbier reaction between aldehydes and ketones with aliphatic or alicyclic halides (see NdI<sub>2</sub>) [Namy et al. *J Organomet Chem* **328** 81 1987.] It is a one-electron reducing agent [Brendt et al. *Aldrichim Acta* **24** 15 1991.]

**SmI<sub>2</sub> in THF** for reactions is prepared in a dry box, or dry gloves, filled with N<sub>2</sub> (not generally necessary for synthetic purposes, but should be used for Grignard-type conditions, i.e. exclusion of moisture) from samarium powder (3g, 20mmol, or distilled metal or ingot), placed in a flask fitted with a dropping funnel containing 1,2-diiodoethane (22.82g, 10mmol) in THF (250ml) which is added dropwise. The reaction has no induction period if the solvent is pure. [Excess metal can be re-used later.] An intense blue-green solution results with a SmI<sub>2</sub> concentration of 4 x 10<sup>-2</sup> M in THF that can be diluted to 10<sup>-1</sup> M. These solutions can be stored for a long time without loss of Sm<sup>2+</sup> concentration if kept in an inert atmosphere in the presence of a small amount of metal. However, it is best to use them within a few days to ensure reproducible results. In addition to promoting coupling reactions, it is also a powerful and selective reducing agent for olefins [Girard et al. *J Am Chem Soc* **102** 2693 1980, see also Asano et al. *Synthesis* 1309 2007], promotes the reaction of chlorofluoroacetonitrile [359-05-7] with aldehydes to form cyanofluorohydrins in the presence of HMPA [Asano et al. *Synthesis* 1309 2007], intramolecular cyclisation reactions [Molander & McKie *J Org Chem* **59** 3186 1994], and intermolecular ketone-olefin coupling reactions [Kawatsura et al. *J Org Chem* **59** 6900 1994].

**Selenious acid** [7783-00-8] M 129.0, m 70°(dec), d<sub>D</sub><sup>20</sup> 3.0, pK<sub>1</sub><sup>25</sup> 2.62, pK<sub>2</sub><sup>25</sup> 8.32 (H<sub>2</sub>SeO<sub>3</sub>). Recrystallise the acid from water. On heating it loses water and SeO<sub>2</sub> sublimes. [Waitkins & Clark *Chem Rev* **36** 235 1945.]

**Selenium** [7782-49-2] M 79.0, m 217.4°, d<sub>4</sub><sup>20</sup> 4.81. Dissolve selenium in small portions in hot conc HNO<sub>3</sub> (2ml/g), filter and evaporate to dryness to give selenious acid which is then dissolved in conc HCl. Pass SO<sub>2</sub> gas through the solution whereby selenium (but not tellurium) precipitates. It is filtered off and washed with conc HCl. This purification process is repeated. The selenium is then converted twice to the selenocyanate by treating with a 10% excess of 3M aqueous KCN (CARE), heated for half an hour on a sand-bath and filtered. Add an equal weight of crushed ice to the cold solution, followed by an excess of cold, conc HCl, with stirring (in an efficient fume cupboard as HCN is evolved) which precipitates selenium powder. This is washed with water until colourless, and then with MeOH and is heated in an oven at 105°. Finally it is fused for 2 hours *in vacuo*. It is cooled, crushed and stored in a desiccator [Tideswell & McCullough *J Am Chem Soc* **78** 3036 1956].

**Selenium dioxide** [7446-08-4] M 111.0, m 340°. Purify it by sublimation at 315°, or by solution in HNO<sub>3</sub>, precipitation of selenium which, after standing for several hours or boiling, is filtered off, then re-oxidised by HNO<sub>3</sub> and cautiously evaporated to dryness below 200°. The dioxide is dissolved in H<sub>2</sub>O and again evaporated to dryness. In H<sub>2</sub>O it forms selenious acid (see selenious acid above). Its solubility in H<sub>2</sub>O is 70%w/w at 20°, and it is soluble in EtOH. It is a useful and selective oxidant. [Waitkins & Clark *Chem Rev* **36** 235 1945, Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed Brauer) Academic Press Vol I p 421 1963.]

**Silica** [7631-86-9 (*colloidal*), 60676-86-0 (*Quartz, Cristobalite, sand*), 112945-52-5 (*fumed*)]. Purification of silica for high technology applications uses isopiestic vapour distillation from concentrated volatile acids and is absorbed in high purity water. The impurities remain behind. Preliminary cleaning to remove surface contaminants uses dip etching in HF or a mixture of HCl, H<sub>2</sub>O<sub>2</sub> and deionised water [Phelan & Powell *Analyst* **109** 1299 1984].

**Silica gel** [63231-67-4, 112926-00-8]. Before use as a drying agent, silica gel is heated in an oven, then cooled in a desiccator. Conditions in the literature range from heating at 110° for 15 hours to 250° for 2-3 hours. Silica gel has been purified by washing with hot acid (in one case successively with *aqua regia*, conc HNO<sub>3</sub>, then conc HCl; in another case it was digested overnight with hot conc H<sub>2</sub>SO<sub>4</sub>), followed by exhaustive washing with distilled water (one week in a Soxhlet apparatus has also been used), and prolonged oven drying. *Alternatively*, silica gel has been extracted with acetone until all soluble material was removed, then dried in a current of air, washed with distilled water and oven dried. Silica gel has also been washed successively with water, M HCl, water, and acetone, then activated at 110° for 15 hours.

**Silicon monoxide** [10097-28-6] **M 44.1, m > 1700°, d<sub>4</sub><sup>20</sup> 2.18**. Purify the monoxide by sublimation in a porcelain tube in a furnace at 1250° (4 hours) in a high vacuum (10<sup>-4</sup>mm) in a stream of N<sub>2</sub>. It is obtained as brownish black scales. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 696 1963.]

**Silicon tetrachloride** [10026-04-7] **M 169.9, m -70°, b 57.6°, d<sub>4</sub><sup>20</sup> 1.483**. Distil it under vacuum and store it in sealed ampoules under N<sub>2</sub>. It fumes in moist air and is very sensitive to moisture. It is soluble in organic solvents. It is a **strong irritant**. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 682-683 1963.]

**12-Silicotungstic acid (tungstosilicic acid; H<sub>4</sub>SiW<sub>12</sub>O<sub>40</sub>)** [12027-43-9] **M 2914.5**. Extract the acid with diethyl ether from a solution acidified with HCl. The diethyl ether is evaporated under vacuum, and the free acid is crystallised twice [Matijevic & Kerker *J Phys Chem* **62** 1271 1958].

**Silver (metal)** [7440-22-4] **M 107.9, m 961.9°, b 2212°, d<sub>4</sub><sup>20</sup> 10.5**. For purification by electrolysis, see Craig et al. [*J Res Nat Bur Stand* **64A** 381 1960]. For purification of crude, or silver residues to pure silver see Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 1028-1030 1963, and for the preparation of *colloidal silver* see *ibid* (Ed. Brauer) p 1034.

**Silver acetate** [563-63-3] **M 166.9, dec on heating, d<sub>25</sub> 3.259**. The salt is a white (or slightly gray if exposed to light) crystalline powder or lustrous needles, which should not be exposed to light for long periods. It can be purified by recrystallisation from H<sub>2</sub>O (preferably containing a little AcOH) as its solubility (w/v) in cold H<sub>2</sub>O is 1% and in boiling H<sub>2</sub>O it is 35%. *Alternatively*, it is boiled with Ac<sub>2</sub>O, the insoluble salt is collected and recrystallised from glacial AcOH. It is freely soluble in dilute HNO<sub>3</sub>. Store in dark bottles away from light. [Beilstein **2 H** 110, **2 I** 48, **2 II** 116, **2 III** 189, **2 IV** 112.]

It oxidises a green solution of U(IV) acetate in liquid NH<sub>3</sub> to the yellow U(VI) in a few hours with the formation of the tris complex (NH<sub>4</sub>)U(VI)O<sub>2</sub>(AcO)<sub>3</sub> [isolated], acetamide and Ag metal [Kline & Kershner *Inorg Chem* **5** 932 1966].

At a 0.2-2 mole%, AgOAc catalyses efficient cycloaddition reactions of methyl isocyanoacetate with olefin Michael-acceptors at ambient temperature to produce Δ<sup>1</sup>- or Δ<sup>2</sup>-pyrrolines in good yields. Isocyanoacetates undergo AgOAc catalysed cyclodimerisation to imidazoles in excellent yields in the absence of a suitable olefin. Also, when combined with azomethine ylide, AgOAc catalyses a 1,3-dipolar cycloaddition reaction in a one-pot sequential cascade process to yield the 7-azabicyclo[2.2.1]heptane ring system which is characteristic of the naturally occurring potent non-opioid analgesic epibatidine [originally isolated from the skin of the Ecuadoran poison frog *Epipedobates tricolor*, Spande et al. *J Am Chem Soc* **114** 3475 1992]. [Grigg et al. *Tetrahedron* **55** 2025 1999.]

When combined with I<sub>2</sub>, AgOAc in aqueous or anhydrous AcOH reacts with olefins to form *vicinal diols* (glycols). The hydroxylations are not always *cis* [Woodward & Brucher *J Am Chem Soc* **80** 209 1958, Ginsberg *J Am Chem Soc* **75** 5746 1953, Barkley et al. *J Am Chem Soc* **76** 5014 1954, Klass et al. *J Am Chem Soc* **77** 2329

1955, Slates & Wendler *J Am Chem Soc* **78** 3749 1956, Jefferies & Milligan *J Chem Soc* 2363 1956, Gunstone & Morris *J Chem Soc* 487 1957, Ellington et al. *J Chem Soc* 1327 1966], but can produce *cis/trans* diol mixtures depending on the olefin [Bunton & Carr *J Chem Soc* 770 1963].

**Silver bromate** [7783-89-3] **M 235.8, m dec on heating,  $d_4^{20}$  5.21**. It crystallises from hot water (80ml/g). It reacts with bromine water to form bromic acid (HBrO<sub>3</sub>) which is a strong oxidising agent. Store it in the dark.

**Silver bromide** [7785-23-1] **M 187.8, m 432°,  $d_4^{20}$  6.47**. Purify it from Fe, Mn, Ni and Zn by zone melting in a quartz vessel under vacuum. It is insoluble in dilute HNO<sub>3</sub> or dilute NH<sub>3</sub> but is soluble in conc NH<sub>3</sub>. Store it in the dark.

**Silver chlorate** [7783-92-8] **M 191.3, m 230°, b 270°(dec),  $d_4^{20}$  4.43**. Recrystallise the chlorate three times from water (10ml/g at 15°; 2ml/g at 80°). Store it in the dark. [Nicholson & Holley *Inorg Synth II* 4 1946, Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 1037 1963.]

**Silver chloride** [7783-90-6] **M 143.3, m 455°, b 1550°,  $d_4^{20}$  5.56**. Recrystallise it from conc NH<sub>3</sub> solution by acidifying with HCl, filtering off the solid, washing it with H<sub>2</sub>O and drying it in a vacuum. It is soluble in NH<sub>3</sub> and should be kept in the dark.

**Silver chromate** [7784-01-2] **M 331.8,  $d_4^{25}$  5.625,  $pK_1^{25}$  0.74,  $pK_2^{25}$  6.49 (for H<sub>2</sub>CrO<sub>4</sub>)**. Wash the red-brown powder with H<sub>2</sub>O, dry it in a vacuum, then powder well and dry again in a vacuum at 90°/5 hours. Its solubility in H<sub>2</sub>O is 0.0014% at 10°. Store it in the dark. [Cardillo & Shimizu *J Org Chem* **42** 4268 1977.]

**Silver cyanide** [506-64-9] **M 133.9, m dec at 320°,  $d_4^{20}$  3.95**. It is a **POISONOUS** white or grayish white powder. Stir it thoroughly with H<sub>2</sub>O, filter, wash well with EtOH and dry it in air in the dark. It is very insoluble in H<sub>2</sub>O (0.000023g in 100ml H<sub>2</sub>O), but is soluble in HCN or aqueous KCN to form the soluble Ag(CN)<sub>2</sub><sup>-</sup> complex. [Schnitz-Dumont *Chem Ber* **72** 298 1939, Randall & Halford *J Am Chem Soc* **52** 184 1930.]

**Silver difluoride** [7783-95-1] **M 145.9, m 690°,  $d_4^{20}$  4.7**. It is highly **TOXIC** because it liberates HF and F<sub>2</sub>. It is very *hygroscopic* and reacts violently with H<sub>2</sub>O. It is a powerful oxidising agent and liberates O<sub>3</sub> from dilute acids, and I<sub>2</sub> from I<sup>-</sup> solution. Store it in quartz or iron ampoules. It is white when pure; otherwise it is brown-tinged. It is thermally stable up to 700°. [Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 241 1963.]

**Silver fluoride** [7775-41-9] **M 126.9, m 435°, b ca 1150°,  $d_4^{20}$  5.852**. The fluoride is a *hygroscopic* solid with a solubility of 135g/100ml of H<sub>2</sub>O at 15°, and forms an insoluble basic fluoride in moist air. Purify it by washing with AcOH and dry \*C<sub>6</sub>H<sub>6</sub>, then keep it in a vacuum desiccator at room temperature to remove \*C<sub>6</sub>H<sub>6</sub>, and store it in opaque glass bottles. The flaky *hygroscopic* crystals darken on exposure to light. It *attacks* bone and teeth. [Sharpe *J Chem Soc* 4538 1952, Kwasnik in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol I p 240 1963.]

**Silver iodate** [7783-97-3] **M 282.8, m >200°,  $d_4^{20}$  5.53**. Wash the iodate with warm dilute HNO<sub>3</sub>, then H<sub>2</sub>O and dry it at 100°, or recrystallise it from NH<sub>3</sub> solution by adding HNO<sub>3</sub>, filtering, washing with H<sub>2</sub>O and drying at 100°.

**Silver nitrate** [7761-88-8] **M 169.9, m 212°, b 444°(dec),  $d_4^{20}$  4.35**. Purify it by recrystallisation from hot water (solubility of AgNO<sub>3</sub> in water is 992g/100ml at 100° and 122g/100ml at 0°). It has also been purified by crystallisation from hot conductivity water by slow addition of freshly distilled EtOH.

**CAUTION: avoid using EtOH for washing the precipitate; and avoid concentrating the filtrate to obtain further crops of AgNO<sub>3</sub> owing to the risk of EXPLOSION (as has been reported to us) caused by the presence of silver fulminate. When using EtOH in the purification, the apparatus should be enveloped in a strong protective shield.** [Tully, *News Ed (Am Chem Soc)* **19** 3092 1941; Garin & Henderson *J Chem Educ*

47 741 1970, Bretherick, *Handbook of Reactive Chemical Hazards* 4th edn, Butterworths, London, 1985, pp 13-14.] Before being used as a standard in volumetric analysis, analytical reagent grade  $\text{AgNO}_3$  should be finely powdered, dried at  $120^\circ$  for 2 hours, then cooled in a desiccator.

Recovery of silver residues as  $\text{AgNO}_3$  [use protective shield during the whole of this procedure] can be achieved by washing with hot water and adding 16M  $\text{HNO}_3$  to dissolve the solid. Filter this through glass wool and concentrate the filtrate on a steam bath until precipitation commences. Cool the solution in an ice-bath and filter the precipitated  $\text{AgNO}_3$ . Dry it at  $120^\circ$  for 2 hours, then cool it in a desiccator in a vacuum. Store it over  $\text{P}_2\text{O}_5$  in a vacuum in the dark. *AVOID contact with hands due to formation of black stains.*

**Silver nitrite** [7783-99-5] **M 153.9, m  $141^\circ(\text{dec})$ ,  $d_4^{20}$  4.45.** Crystallise the salt from hot conductivity water (70ml/g) in the dark. Dry it in the dark under vacuum. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 1048 1963.]

**Silver(I) oxide** [20667-12-3] **M 231.7, m  $\sim 200^\circ(\text{dec})$ ,  $d_4^{20}$  7.13.** Leach the oxide with hot water in a Soxhlet apparatus for several hours to remove any entrained electrolytes. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1037 1965.]

**Silver (II) oxide** [1301-96-8] **M 123.9, m  $>100^\circ(\text{dec})$ ,  $d^{25}$  7.22.** It is soluble in 40,000 parts of  $\text{H}_2\text{O}$ , and should be protected from light. Stir it with an alkaline solution of potassium peroxy sulfate ( $\text{K}_2\text{S}_2\text{O}_8$ ) at  $85\text{--}90^\circ$ . The black  $\text{AgO}$  is collected, washed free from sulfate with  $\text{H}_2\text{O}$  made slightly alkaline and dried in air in the dark. [Hammer & Kleinberg *Inorg Synth* IV 12 1953.]

**Silver perchlorate monohydrate** [14242-05-8 ( $\text{H}_2\text{O}$ ), 7783-93-9 (anhydrous)] **M 207.3,  $\text{pK}^{25}$  -2.4 to -3.1 (for  $\text{HClO}_4$ ).** Reflux it with \*benzene (6ml/g) in a flask fitted with a Dean and Stark trap until all the water is removed azeotropically (ca 4 hours). The solution is cooled and diluted with dry pentane (4ml/g of  $\text{AgClO}_4$ ). The precipitated  $\text{AgClO}_4$  is filtered off and dried in a desiccator over  $\text{P}_2\text{O}_5$  at 1mm for 24 hours [Radell et al. *J Am Chem Soc* 83 3958 1961]. It has also been recrystallised from perchloric acid. [**Caution due to its EXPLOSIVE nature in the presence of organic matter.**] Store it in the dark.

**Silver permanganate** [7783-98-4] **M 226.8,  $d_4^{20}$  4.49.** The salt forms violet crystals which can be crystallised from hot  $\text{H}_2\text{O}$  (soluble is 9g/L at  $20^\circ$ ). Store it in the dark. This oxidising agent is decomposed by light. [Lux in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 1463 1963.]

**Silver sulfate** [10294-26-5] **M 311.8, m  $652^\circ$ , b  $1085^\circ(\text{dec})$ ,  $d_4^{20}$  5.45.** Crystallise the sulfate from hot conc  $\text{H}_2\text{SO}_4$  containing a trace of  $\text{HNO}_3$ , and dilute with  $\text{H}_2\text{O}$  while being strongly cooled. The precipitate is filtered off, washed with  $\text{H}_2\text{O}$  and dried at  $120^\circ$ . Its solubility in  $\text{H}_2\text{O}$  is 0.8% at  $17^\circ$ , and 1.46% at  $100^\circ$ . Store it in the dark. [Glemser & Sauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1042 1965.]

**Silver tetraoxorhenate** [ $\text{Ag}(\text{ReO}_4)$ ] [7784-00-1] **M 358.1,  $\text{pK}_a$  -1.25.** It is prepared by adding slowly, with ice cooling,  $\text{H}_2\text{O}_2$  (35%, 100ml) to a stirred suspension of Re powder (10g, 54mmol) in  $\text{H}_2\text{O}$  (20ml) during 4 hours, stirring at  $\sim 25^\circ$  for 30 minutes, and then at  $80^\circ$  for 3 hours. The mixture is filtered from a small amount of insoluble material, and the clear solution is treated with  $\text{Ag}(\text{NO}_3)$  (10g, 59mmol), when  $\text{Ag}(\text{ReO}_4)$  immediately separates as a white precipitate which is filtered off, washed with  $\text{Et}_2\text{O}$  (3 x 25ml) to remove  $\text{H}_2\text{O}$  and  $\text{H}_2\text{O}_2$ , and dried *in vacuo* to give the silver salt (18.7g, 97%). [Herrmann & Kratzer *Inorg Synth* 33 111 2002, Herrmann et al. *Angew Chem, Int Ed. Engl* 36 2652 1997.]

**Silver thiocyanate** [1701-93-5] **M 165.9, m  $265^\circ(\text{dec})$ ,  $d_4^{20}$  3.746,  $\text{pK}^{25}$  -1.85 (for HSCN).** Digest the solid salt with dilute aqueous  $\text{NH}_4\text{NCS}$ , filter, wash it thoroughly with  $\text{H}_2\text{O}$  and dry it at  $110^\circ$  in the dark. It is soluble in dilute aqueous  $\text{NH}_3$ . *Alternatively*, dissolve it in strong aqueous  $\text{NH}_4\text{NCS}$  solution, filter and dilute with large volumes of  $\text{H}_2\text{O}$  when the Ag salt separates. The solid is washed with  $\text{H}_2\text{O}$  by decantation until free from  $\text{NCS}^-$  ions, collected, washed with  $\text{H}_2\text{O}$ ,  $\text{EtOH}$  and dried in an air oven at  $120^\circ$ . It has also been purified by dissolving in dilute aqueous  $\text{NH}_3$  when single crystals are formed by free evaporation of the solution in air. Store it in the

dark. [Garrick & Wilson *J Chem Soc* 835 1932, Occlshaw *J Chem Soc* 2405 1932, IR and Raman: *Acta Chem Scand* 13 1607 1957, Lindqvist *Acta Cryst* 10 29 1957.]

**Sodium (metal)** [7440-23-5] **M 23.0, m 97.5°, d<sub>4</sub><sup>20</sup> 0.97.** The metal is placed on a coarse grade of sintered-glass filter, melted under vacuum and forced through the filter using argon. The Pyrex apparatus is then re-evacuated and sealed off below the filter, so that the sodium could be distilled at 460° through a side arm and condenser into a receiver bulb which is then sealed off [Gunn & Green *J Am Chem Soc* 80 4782 1958]. **EXPLODES and IGNITES in water.**

**Sodium amide** [7782-92-5] **M 39.0, m 210°.** It reacts *violently* with H<sub>2</sub>O and is soluble in liquid NH<sub>3</sub> (1% at 20°). It should be stored in wax-sealed containers in small batches. It is very *hygroscopic* and absorbs CO<sub>2</sub> and H<sub>2</sub>O. If the solid is discoloured by being yellow or brown in colour, then it should be destroyed as it can be highly **EXPLOSIVE**. It should be replaced if discoloured. It is best destroyed by covering it with much toluene and slowly adding dilute EtOH with stirring until all the ammonia is liberated (FUME CUPBOARD). [Dennis & Bourne *Inorg Synth* I 74 1939, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 465 1963, Bergstrom *Org Synth Coll Vol III* 778 1955.]

**Sodium ammonium hydrogen phosphate** [13011-54-6] **M 209.1, m 79°(dec), d<sub>4</sub><sup>20</sup> 1.55.** Crystallise it from hot water (1ml/g).

**Sodium arsenate heptahydrate** [10048-95-0] **M 312.0, m 50 (loses 5H<sub>2</sub>O), m 130°, d<sub>4</sub><sup>20</sup> 1.88 pK<sub>1</sub><sup>25</sup> 2.22, pK<sub>2</sub><sup>25</sup> 6.98 (for H<sub>3</sub>AsO<sub>4</sub>).** Crystallise it from water (2ml/g).

**Sodium azide** [26628-22-8] **M 65.0, m 300°(dec, explosive), pK<sub>25</sub> 4.72 (for HN<sub>3</sub>).** Crystallise sodium azide from hot water or from water by adding absolute EtOH or acetone. Also purify it by repeated crystallisation from an aqueous solution saturated at 90° by cooling it to 10°, and adding an equal volume of EtOH. The crystals are washed with acetone, and the azide is dried at room temperature under vacuum for several hours in an Abderhalden pistol. Its solubility in H<sub>2</sub>O is 42% at 18°, and in EtOH it is 0.22% at 0°. [Das et al. *J Chem Soc, Faraday Trans 1* 78 3485 1982, Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 474-475 1963, Browne *Inorg Synth* 1 79 1939, Frierson *Inorg Synth* II 139 1946.] **HIGHLY POISONOUS and potentially explosive.**

**Sodium bicarbonate** [144-55-8] **M 84.0, m ~50°(dec, -CO<sub>2</sub>).** Crystallise it from hot water (6ml/g). The solid should not be heated above 40° due to the formation of carbonate.

**Sodium bisulfite** [7631-90-5] **M 104.1, d<sub>4</sub><sup>20</sup> 1.48.** Crystallise it from hot H<sub>2</sub>O (1ml/g). Dry it at 100° under vacuum for 4 hours.

**Sodium borate (borax)** [1330-43-4] **M 201.2, m 741°, d<sub>4</sub><sup>20</sup> 2.37.** Most of the water of hydration is removed from the *decahydrate* (see below) by evacuation at 25° for three days, followed by heating to 100° and evacuation with a high-speed diffusion pump. The dried sample is then heated gradually to fusion (above 966°), allowed to cool gradually to 200°, then transferred to a desiccator containing P<sub>2</sub>O<sub>5</sub> [Grenier & Westrum *J Am Chem Soc* 78 6226 1956]. [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 794-795 1963.]

**Sodium borate (decahydrate, hydrated borax)** [1303-96-4] **M 381.2, m 75°(loses 5H<sub>2</sub>O at 60°), d<sub>4</sub><sup>20</sup> 1.73.** Crystallise the borate from water (3.3ml/g), keeping below 55° to avoid formation of the pentahydrate. Filter it off at the pump, wash it with water and equilibrate it for several days in a desiccator containing an aqueous solution saturated with respect to sucrose and NaCl. Borax can be prepared more quickly (but its water content is somewhat variable) by washing the recrystallised material at the pump with water, followed by 95% EtOH, then Et<sub>2</sub>O, and dried in air at room temperature for 12-18 hours on a clock glass. [Becher in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 794-795 1963.]

**Sodium borohydride** [16940-66-2] **M 37.8, m ~400°(dec), d<sub>4</sub><sup>20</sup> 1.07.** After adding NaBH<sub>4</sub> (10g) to freshly distilled diglyme (120ml) in a dry three-necked flask fitted with a stirrer, nitrogen inlet and outlet, the mixture is

stirred for 30 minutes at 50° until almost all of the solid has dissolved. Stirring is stopped, and, after the solid has settled, the supernatant liquid is forced under N<sub>2</sub> pressure through a sintered-glass filter into a dry flask. [The residue is centrifuged to obtain more of the solution which is added to the bulk.] The solution is cooled slowly to 0° and then decanted from the white needles that separated. The crystals are dried by evacuating for 4 hours to give anhydrous NaBH<sub>4</sub>. *Alternatively*, after the filtration at 50° the solution is heated at 80° for 2 hours to give a white precipitate of substantially anhydrous NaBH<sub>4</sub> which is collected on a sintered-glass filter under N<sub>2</sub>, then evacuated at 60° for 2 hours [Brown et al. *J Am Chem Soc* 77 6209 1955].

NaBH<sub>4</sub> has also been crystallised from isopropylamine by dissolving it in the solvent at reflux, cooling, filtering and allowing the solution to stand in a filter flask connected to a Dry-ice/acetone trap. After most of the solvent has passed over into the cold trap, crystals are removed with forceps, washed with dry diethyl ether and dried under vacuum. [Kim & Itoh *J Phys Chem* 91 126 1987.] Somewhat less pure crystals were obtained more rapidly by using Soxhlet extraction with only a small amount of solvent and extracting for about 8 hours. The crystals that formed in the flask are filtered off, then washed and dried as before. [Stockmayer et al. *J Am Chem Soc* 77 1980 1955.] Other solvents used for crystallisation include water and liquid ammonia.

**Sodium bromate** [7789-38-0] M 150.9, m 381°, d<sub>4</sub><sup>20</sup> 3.3. It is crystallised from hot water (1.1ml/g) to decrease contamination by NaBr, bromine and hypobromite. [Noszticzius et al. *J Am Chem Soc* 107 2314 1985.]

**Sodium bromide** [7647-15-6] M 102.9, m 747°, b 1390°, d<sub>4</sub><sup>20</sup> 3.2. Crystallise the bromide from water (0.86ml/g) between 50° and 0°, and dry it at 140° under vacuum (this purification may not eliminate chloride ion).

**Sodium carbonate** [497-19-8] M 106.0, m 858°, d<sub>4</sub><sup>20</sup> 2.5. It crystallises from water as the *decahydrate* which is redissolved in water to give a near-saturated solution. By bubbling CO<sub>2</sub>, NaHCO<sub>3</sub> is precipitated. It is filtered off, washed and ignited for 2 hours at 280° [MacLaren & Swinehart *J Am Chem Soc* 73 1822 1951]. Before being used as a volumetric standard, analytical grade material should be dried by heating at 260-270° for 0.5 hour and allowed to cool in a desiccator. It has a transition point at 450°, and its solubility in water is 21.58% at 20° (*decahydrate* in solid phase), 49.25% at 35° (*heptahydrate* in solid phase) and 44.88% at 75° (*monohydrate* in solid phase) [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 987-988 1963]. After three recrystallisations, technical grade Na<sub>2</sub>CO<sub>3</sub> had Cr, Mg, K, P, Al, W, Sc and Ti at 32, 9.4, 6.6, 3.6, 2.4, 0.6, 0.2 and 0.2 ppm respectively; another technical source had Cr, Mg, Mo, P, Si, Sn and Ti at 2.6, 0.4, 4.2, 13.4, 32, 0.6, 0.8 ppm respectively.

**Sodium chlorate** [7775-09-9] M 106.4, m 248°, b >300°(dec), d<sub>4</sub><sup>20</sup> 2.5. It is crystallised from hot water (0.5ml/g). It is a strong oxidising agent, and should be kept clear from organic matter.

**Sodium chloride** [7647-14-5] M 58.4, m 800.7°, b 1413°, d<sub>4</sub><sup>20</sup> 2.17. It is recrystallised from a saturated aqueous solution (2.7ml/g) by passing in HCl gas, or by adding EtOH or acetone. It can be freed from bromide and iodide impurities by adding chlorine water to an aqueous solution and boiling it for some time to expel free bromine and iodine. Traces of iron can be removed by prolonged boiling of solid NaCl in 6M HCl; the crystals are then washed with EtOH and dried at *ca* 100°. Sodium chloride has been purified by sublimation in a stream of pre-purified N<sub>2</sub> and collected by electrostatic discharge [Ross & Winkler *J Am Chem Soc* 76 2637 1954]. For use as a primary analytical standard, analytical reagent grade NaCl should be finely ground, dried in an electric furnace at 500-600° in a platinum crucible, and allowed to cool in a desiccator. For most purposes, however, drying at 110-120° is satisfactory.

**Sodium chlorite** [7758-19-2] M 90.4, m ~180°(dec). Crystallise the chlorite from hot water and store it in a cool place. It has also been crystallised from MeOH by counter-current extraction with liquid ammonia [Curti & Locchi *Anal Chem* 29 534 1957]. A major impurity is chloride ion which can be removed by recrystallisation from 0.001M NaOH. [Schmeisser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 312 1963.]

**Sodium chromate tetrahydrate** [10034-82-9] M 234.0, m ~20°(for 10H<sub>2</sub>O), d<sub>4</sub><sup>20</sup> 2.7, pK<sub>1</sub><sup>25</sup> 0.74, pK<sub>2</sub><sup>25</sup> 6.49 (for H<sub>2</sub>CrO<sub>4</sub>). Crystallise the chromate from hot water (0.8ml/g). It is *deliquescent*.

**Sodium cyanate** [917-61-3] **M 65.0, m 550°, d<sub>4</sub><sup>20</sup> 1.893, pK<sup>25</sup> 3.47 (for HCNO).** It forms colourless needles from EtOH. Its solubility in EtOH is 0.22g/100g at 0°C. It is soluble in H<sub>2</sub>O but can be recrystallised from small volumes of it.

**Sodium cyanoborohydride** [25895-60-7] **M 62.8, m 240-242°(dec), d<sup>28</sup> 1.20.** It is a very *hygroscopic* solid, soluble in H<sub>2</sub>O (212% at 29°, 121% at 88°), tetrahydrofuran (37% at 28°, 42.2% at 62°), very soluble in MeOH, slightly soluble in EtOH, but insoluble in Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> and hexane. It is stable to acid up to pH 3 but is hydrolysed in 12N HCl. The rate of hydrolysis at pH 3 is 10<sup>-8</sup> times that of NaBH<sub>4</sub>. The fresh commercially available material is usually sufficiently pure. If very pure material is required, one of the following procedures can be used [Lane *Synthesis* 135 1975]: (a) The NaBH<sub>3</sub>CN is dissolved in tetrahydrofuran (20% w/v), filtered and the filtrate is treated with a fourfold volume of CH<sub>2</sub>Cl<sub>2</sub>. The solid is collected and dried in a vacuum [Wade et al. *Inorg Chem.* 9 2146 1970]. (b) NaBH<sub>3</sub>CN is dissolved in dry MeNO<sub>2</sub>, filtered, and the filtrate is poured into a 10-fold volume of CCl<sub>4</sub> with vigorous stirring; the white precipitate is collected, washed several times with CCl<sub>4</sub> and dried in a vacuum (yield 75%) [Berschied & Purcell *Inorg Chem* 9 624 1970]. (c) When the above procedures fail to give a clean product, then dissolve NaBH<sub>3</sub>CN (10g) in tetrahydrofuran (80ml) and add N MeOH/HCl until the pH is 9. Pour the solution with stirring into dioxane (250ml). The solution is filtered and heated to reflux. A further volume of dioxane (150ml) is added slowly with swirling. The solution is cooled slowly to room temperature, then chilled in ice and the crystalline dioxane complex is collected, dried in a vacuum for 4 hours at 25°, then 4 hours at 80° to yield the amorphous dioxane-free powder (6.7g) with purity >98% [Borch et al. *J Am Chem Soc* 93 2897 1971]. The purity can be checked by iodometric titration [Lyttle et al. *Anal Chem* 24 1843 1952].

**Sodium dichromate** [7789-12-0] **M 298.0, m 84.6° (2H<sub>2</sub>O), 356° (anhydrous); b 400°(dec), d<sub>4</sub><sup>25</sup> 2.348.** Crystallise the dichromate from small volumes of H<sub>2</sub>O by evaporation to crystallisation. Its solubility in H<sub>2</sub>O is 238% at 0° and 508% at boiling. The red *dihydrate* is slowly dehydrated by heating at 100° for long periods. It is *deliquescent* and is a powerful oxidising agent—*do not place it in contact with skin—wash immediately as it is caustic.* (Possible **carcinogen**.)

**Sodium dihydrogen orthophosphate dihydrate** [13472-35-0 (2H<sub>2</sub>O), 10049-21-5 (H<sub>2</sub>O), 7558-80-7 (anhydrous)] **M 156.0, m 60°(dec), d<sub>4</sub><sup>20</sup> 1.91.** Crystallise it from warm water (0.5ml/g) by chilling.

**Sodium dithionite dihydrate** [7631-94-9] **M 242.1, m 110°(loses 2H<sub>2</sub>O), 267°(dec), d<sub>4</sub><sup>20</sup> 2.19, pK<sub>Est(1)</sub><sup>25</sup> -3.4, pK<sub>2</sub><sup>25</sup> 0.49 (for dithionic acid).** Crystallise it from hot water (1.1ml/g) by cooling.

**Sodium ferricyanide monohydrate** [14217-21-1, 13601-19-9 (anhydrous)] **M 298.9, pK<sup>25</sup> <1 (for ferricyanide).** Crystallise the ferricyanide from hot water (1.5ml/g) or by precipitation from 95% EtOH.

**Sodium ferrocyanide decahydrate** [13601-19-9] **M 484.1, m 50-80° (loses 10H<sub>2</sub>O), 435°(dec), d<sub>4</sub><sup>20</sup> 1.46, pK<sub>3</sub><sup>25</sup> 2.57, pK<sub>4</sub><sup>25</sup> 4.35 (for ferrocyanide).** Crystallise it from hot water (0.7ml/g), until free of ferricyanide as shown by the absence of formation of Prussian Blue colour with ferrous sulfate solution.

**Sodium fluoride** [7681-49-4] **M 42.0, m 996°, b 1695°, d<sub>4</sub><sup>20</sup> 2.56.** Crystallise NaF from water by partial evaporation in a vacuum desiccator, or dissolve it in water, and precipitate *ca* half of it by adding EtOH. The precipitate is dried in an air oven at 130° for one day, and then stored in a desiccator over KOH. Its solubility in H<sub>2</sub>O is 4% at 15° and 4.3% at 25°. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 235 1963].

**Sodium fluoroborate** [13755-29-8] **M 109.8, m 384°, d<sub>4</sub><sup>20</sup> 2.47, pK<sup>25</sup> -4.9 (for fluoroboric acid H<sub>3</sub>O<sup>+</sup>BF<sub>4</sub><sup>-</sup>).** Crystallise the fluoroborate from hot water (50ml/g) by cooling to 0°. *Alternatively*, free it from insoluble material by dissolving it in a minimum amount of water; then fluoride ions are removed by adding concentrated lanthanum nitrate in excess. After removing lanthanum fluoride by centrifugation, the supernatant is passed through a cation-exchange column (Dowex 50, Na<sup>+</sup>-form) to remove any remaining lanthanum [Anbar & Guttman *J Phys Chem* 64 1896 1960]. It has also been recrystallised from anhydrous MeOH and dried in a vacuum at 70° for 16 hours. Keep it dry as it is *hygroscopic*. [Delville et al. *J Am Chem Soc* 109 7293 1987.]

**Sodium fluorosilicate** [16893-85-9] **M 188.1**. Crystallise it from hot water (40ml/g) by cooling.

**Sodium hexafluorophosphate** [21324-39-0] **M 167.9**,  $\text{pK}_1^{25} \sim 0.5$ ,  $\text{pK}_2^{25} 5.12$  (for fluoro-phosphoric acid  $\text{H}_2\text{PO}_3\text{F}$ ). Recrystallise it from acetonitrile and dry it in a vacuum for 2 days at room temperature. It is an irritant and is *hygroscopic*. [Delville et al. *J Am Chem Soc* **109** 7293 1987.]

**Sodium hexanitrocobaltate III** ( $\text{Na}_3[\text{Co}(\text{NO})_6]$ ) [13600-98-1] **M 403.9**. Dissolve the salt (*ca* 60g) in  $\text{H}_2\text{O}$  (300ml), filter to obtain a clear solution, add 96% EtOH (250ml) with vigorous stirring. Allow the precipitate to settle for 2 hours, filter it off, wash it with EtOH (4 x 25ml), twice with  $\text{Et}_2\text{O}$  and dry it in air [Glemser in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1541 1965]. It forms yellow to brown crystals which are very soluble in  $\text{H}_2\text{O}$ , are decomposed by acid and form an insoluble K salt. Used for estimating potassium.

**Sodium hydride (NaH)** [7646-69-7] **M 24.00**,  $m \sim 425^\circ(\text{dec})$ ,  $800^\circ(\text{dec})$ , **d 1.396**. NaH is a granular gray powder which ignites in moist air and should be stored in an inert atmosphere. It is prepared in the laboratory by passing  $\text{H}_2$  through a dispersion of Na in an industrial white oil at *ca*  $250^\circ$  [Mattson & Whaley *Inorg Synth* **5** 10 1957]. Like sodium it is very reactive with  $\text{H}_2\text{O}$  liberating hydrogen and NaOH. Fortunately it is commercially available as a 50 to 60% (w/w) dispersion in oil, which protects the NaH from the atmosphere, moderates its reaction with moisture, and can be handled routinely without recourse to using dry box techniques. In this form its reaction with  $\text{H}_2\text{O}$  (and alcohols) can be controlled safely, unlike the dry solid. The oil is freely soluble in petroleum ethers, hexanes,  $\text{Et}_2\text{O}$ ,  $^*\text{C}_6\text{H}_6$ , toluene and THF without affecting NaH reactivity. Removal of the oil is achieved by adding any of these solvents, decanting them off, repeating the process, then adding the desired reaction solvent to cover the NaH. NaH is insoluble in inert organic solvents and in liquid  $\text{NH}_3$ . When the dispersion is kept for long periods, or if the lid of the container in which it is kept is removed a considerable number of times, the NaH may have deteriorated and the mixture will contain a high percentage of NaOH and/or  $\text{NaHCO}_3$  or  $\text{Na}_2\text{CO}_3$ . In this case a new batch should be used as laboratory purification is not practicable. The NaH content can, however, be determined by carefully adding  $\text{H}_2\text{O}$  to a known weight of the dispersion oil in a flask and measuring the volume of  $\text{H}_2$  released using a manometer. The aqueous solution can then be titrated with standard acid (phenolphthalein indicator) to provide the total Na content. [For numerous examples of the use of NaH in organic synthesis see the various volumes of Fieser and Fieser's *Reagents in Organic Synthesis*.]

**Sodium hydroxide (anhydrous)** [1310-73-2] **M 40.0**,  $m 323^\circ$ ,  $b 1390^\circ$ ,  $d_4^{20} 2.13$ . Common impurities are water and sodium carbonate. Sodium hydroxide can be purified by dissolving 100g in 1L of pure EtOH, filtering the solution under vacuum through a fine sintered-glass disc to remove insoluble carbonates and halides. (This and subsequent operations should be performed in a dry,  $\text{CO}_2$ -free box.) The solution is concentrated under vacuum, using mild heating, to give a thick slurry of the mono-alcoholate which is transferred to a coarse sintered-glass disc and evacuated free of mother liquor. After washing the crystals several times with purified alcohol to remove traces of water, they are dried in a vacuum, with mild heating, for about 30 hours to decompose the alcoholate, leaving a fine white crystalline powder [Kelly & Snyder *J Am Chem Soc* **73** 4114 1951]. **CAUSTIC.**

**Sodium hydroxide solutions (caustic)**,  $\text{pK}^{25} 14.77$ . Carbonate ion can be removed by passage through an anion-exchange column (such as Amberlite IRA-400;  $\text{OH}^-$ -form). The column should be freshly prepared from the chloride form by slow prior passage of sodium hydroxide solution until the effluent gives no test for chloride ions. After use, the column can be regenerated by washing with dilute HCl, then water. Similarly, other metal ions are removed when a 1M (or more dilute) NaOH solution is passed through a column of Dowex ion-exchange A-1 resin in its  $\text{Na}^+$ -form.

*Alternatively*, carbonate contamination can be reduced by rinsing sticks of NaOH (analytical reagent quality) rapidly with  $\text{H}_2\text{O}$ , then dissolving in distilled  $\text{H}_2\text{O}$ , or by preparing a concentrated aqueous solution of NaOH and drawing off the clear supernatant liquid. (Insoluble  $\text{Na}_2\text{CO}_3$  is left behind.) Carbonate contamination can be reduced by adding a slight excess of concentrated  $\text{BaCl}_2$  or  $\text{Ba}(\text{OH})_2$  to a NaOH solution, shaking well and allowing the  $\text{BaCO}_3$  precipitate to settle. If the presence of Ba in the solution is unacceptable, an electrolytic purification can be used. For example, sodium amalgam is prepared by the electrolysis of 3L of 30% NaOH

with 500ml of pure mercury for cathode, and a platinum anode, passing 15 Faradays at 4Amps, in a thick-walled polyethylene bottle. The bottle is then fitted with inlet and outlet tubes, the spent solution being flushed out by CO<sub>2</sub>-free N<sub>2</sub>. The amalgam is then washed thoroughly with a large volume of deionized water (with the electrolysis current switched on to minimise loss of Na). Finally, a clean steel rod is placed in contact in the solution with the amalgam (to facilitate hydrogen evolution), reaction being allowed to proceed until a suitable concentration is reached, before being transferred to a storage vessel and diluted as required [Marsh & Stokes *Aust J Chem* **17** 740 1964].

**Sodium hypophosphite monohydrate** [10039-56-2] **M 106.0** (see pK of hypophosphorous acid). Dissolve it in boiling EtOH, cool and add dry Et<sub>2</sub>O till all the salt separates. Collect and dry it in vacuum. It is soluble in 1 part of H<sub>2</sub>O. It liberates PH<sub>3</sub> on heating and can *ignite* spontaneously when heated. The anhydrous salt is soluble in ethylene glycol (33% w/w) and propylene glycol (9.7%) at 25°.

**Sodium iodate** [7681-55-2] **M 197.9, m dec on heating, d<sub>4</sub><sup>20</sup> 4.28**. Crystallise sodium iodate from water (3ml/g) by cooling.

**Sodium iodide** [7681-82-5] **M 149.9, m 660°, b 1304°, d<sub>4</sub><sup>20</sup> 3.67**. Crystallise NaI from water/ethanol solution and dry it for 12 hours under vacuum, at 70°. *Alternatively*, dissolve it in acetone, filter it and cool it to -20°; the resulting yellow crystals are filtered off and heated in a vacuum oven at 70° for 6 hours to remove acetone. The NaI is then crystallised from very dilute NaOH, dried under vacuum, and stored in a vacuum desiccator [Verdin *Trans Faraday Soc* **57** 484 1961].

**Sodium-mercury amalgam (sodium amalgam)** [11110-52-4] **M depends on Na%, m (Na-Hg/m): 0.5%/0°, 1.0%/50°, 1.5%/100°, 2.0%/130°, 2.5%/156°, 3.3%/250°, 4.0%/320°**. The composition of the amalgam depends on the amount of Na added to the Hg. The reaction is exothermic and can be violent on first mixing. Generally 1-3% of Na in Hg is used, but higher amalgams can be obtained or prepared. Amalgams with more than 1% are solids with melting points that increase with the percentage of Na reaching > 360° at > 5% Na. Unless it has been prepared recently, it is always best to prepare the amalgam afresh. Three general procedures can be adopted. Use a fume cupboard as Hg vapour is **TOXIC**.

*Method 1:* For a ~1.2% amalgam, clean dry Na (9g) is placed in a 500ml flask with a reflux condenser, covered with dry toluene (20ml), and warmed gently until the Na has melted. Hg is added dropwise (separating funnel) with swirling. The exothermic reaction during the addition of the first few mls of Hg will cause the toluene to reflux, and Hg is added at such a rate as to keep the toluene refluxing gently. When addition is complete, allow the solvent to evaporate and the amalgam will solidify. Note that it is low melting (~50°) but will solidify on cooling. The solid amalgam is pulverised in a mortar (under N<sub>2</sub>; to avoid spluttering of small quantities of solid Na-Hg, the mortar is covered with a thick filter paper, or cardboard, with a hole in the centre to allow the pestle to go through), and stored away from moisture in a tightly stoppered container. Up to a ~2%Na amalgam can be prepared by this method and using 15.2g Na and 750ml of Hg.

*Method 2:* To a flask with dry N<sub>2</sub> or Argon flushing continuously through it, clean freshly cut Na (22.8g) is added followed by Hg (10ml) from a separating funnel. The flask is warmed gently until the reaction commences then more Hg (total 750g) is added, with shaking, at such a rate as to keep the amalgam molten. After half the Hg is added, the amalgam begins to solidify and should be kept liquid by external heating. When addition is complete, the molten amalgam is poured (under N<sub>2</sub> or Argon) into a cold mortar and powdered (see Method 1 above). This provides a 3% amalgam. It should be stored as before. A 5%Na amalgam corresponds to a composition of NaHg<sub>2</sub>, and a 20% amalgam corresponds to a composition of Na<sub>2</sub>Hg.

*Method 3:* This method is convenient and commonly used. To the required volume of Hg in a wide-necked flask continuously flushed with N<sub>2</sub> or Argon, is added Na which is cut into small cubes, ca 5mm, and each cube is speared with a pointed glass rod and pushed into the liquid Hg. If traces of metal are not a concern, then it is better to use a stainless steel spike or a long tweezers to pick the Na cubes. A vigorous reaction occurs (perhaps with a flash of light occasionally) as each piece of Na dissolves into the Hg or amalgam. The temperature rises and the amalgam is kept in liquid form by the heat released on adding the Na. Do not allow the amalgam to solidify before all the Na is added. If it does, then heat the flask with a burner to keep the amalgam in a liquid state. To avoid damage from flying pieces of Na resulting from the vigorous reaction, the glass rod or metal pick spearing the Na cubes is inserted through a hole in the centre of a thick cardboard sheet which is placed over the mouth of the flask. The amalgam is then powdered and stored as above.

The amalgams are good reducing agents, and react with H<sub>2</sub>O releasing “nascent H” and H<sub>2</sub>, but much less vigorously than Na metal. [Brauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1802 1965, Audrieth *Inorg Synth* Vol I 5 1939, Babcock *Inorg Synth* I 10 1939, Holleman *Org Synth Coll Vol I* 555 1941, Renfrow & Hauser *Org Synth Coll Vol II* 609 1943, Deulofeu & Guerrero *Org Synth Coll Vol III* 589 1955.]

**Sodium metaperiodate (NaIO<sub>4</sub>)** [7790-28-5] M 213.9, m ~300°(dec), d<sub>4</sub><sup>20</sup> 4.17. Recrystallise it from hot water. [Willard *Inorg Synth* I 170 1939, Bernays *Inorg Synth* II 212 1946.]

**Sodium metasilicate pentahydrate** [6834-92-0] M 212.1, m 1088°, d<sub>4</sub><sup>20</sup> 2.4. Crystallise it from aqueous 5% NaOH solution. [Schwartz *Z Anorg Allgem Chem* 126 62 1923.]

**Sodium molybdate dihydrate** [10102-40-6] M 241.9, m 100°(loses 2H<sub>2</sub>O), 687°, d<sub>4</sub><sup>20</sup> 3.28, pK<sup>25</sup> 4.08 (for H<sub>2</sub>MoO<sub>4</sub>). Crystallise it from hot water (1ml/g) by cooling to 0°.

**Sodium nitrate** [7631-99-4] M 85.0, m 307°, b 380°, d<sub>4</sub><sup>20</sup> 2.26. Crystallise NaNO<sub>3</sub> from hot water (0.6ml/g) by cooling to 0°, or from a concentrated aqueous solution by adding MeOH. Dry it under a vacuum at 140°. After two recrystallisations, technical grade sodium nitrate had K, Mg, B, Fe Al, and Li at 100, 29, 0.6, 0.4, 0.2 and 0.2 ppm respectively. (See KNO<sub>3</sub>.)

**Sodium nitrite** [7632-00-0] M 69.0, m 271°, b 320°, d<sub>4</sub><sup>20</sup> 2.17. Crystallise NaNO<sub>2</sub> from hot water (0.7ml/g) by cooling to 0°, or from its own melt. Dry it over P<sub>2</sub>O<sub>5</sub>. (See KNO<sub>2</sub>.)

**Sodium oleate** [143-19-1] M 304.4, m 233-235°. It is recrystallised from EtOH, and is dried in an oven at 100°. [Beilstein 2 H 465, 2 I 201, 2 II 434, 2 III 1405, 2 IV 1645.]

**Sodium percarbonate (Na<sub>2</sub>CO<sub>3</sub>·1.5H<sub>2</sub>O<sub>2</sub>, hydrogen peroxide sodium carbonate adduct)** [15630-89-4] M 157.0, dec >120°. This is not a per-salt, but is crystalline Na<sub>2</sub>CO<sub>3</sub> with H<sub>2</sub>O<sub>2</sub> of crystallisation with a specific composition. The crystals are consistent in having 1.5 molecules of H<sub>2</sub>O<sub>2</sub> for each molecule of Na<sub>2</sub>CO<sub>3</sub>. It is a stable, inexpensive, a non-toxic source of H<sub>2</sub>O<sub>2</sub> that is easily handled in the laboratory. It is prepared by the addition of excess of H<sub>2</sub>O<sub>2</sub> (~3 to 15%) to a saturated aqueous solution of analytically pure Na<sub>2</sub>CO<sub>3</sub> or NaHCO<sub>3</sub> at room temperature, stirring for 30 minutes then absolute EtOH is added. The percarbonate separates as white needle-shaped crystals. These are filtered off, washed with EtOH then Et<sub>2</sub>O and dried in air. Its available oxygen content is determined by titration with standard KMnO<sub>4</sub> after acidification, and is ~13-15%. [Galwey & Hood *J Phys Chem* 83 1810 1979.] Its solubility in H<sub>2</sub>O at ~25° is ~1 mol/L and the pH is ~10.5. The kinetics of thermal decomposition have been studied in detail and are autocatalytic, and best represented by the equation Na<sub>2</sub>CO<sub>3</sub>·1.5H<sub>2</sub>O<sub>2</sub> = Na<sub>2</sub>CO<sub>3</sub> + 1.5 H<sub>2</sub>O + 0.75O<sub>2</sub>. [Galwey & Hood *J Chem Soc, Faraday Trans* 78 2815 1982.] It is a useful oxidising agent; converting nitriles to amides [Kabalka et al. *Synth Commun* 20 1445 1990], oxidises α-diketones to dicarboxylic acids [Yang et al. *Synth Commun* 23 1183 1993], it promotes the Baeyer-Villiger reaction in TFA at 0° [Olah et al. *Synthesis* 739 1991], forms epoxides with olefins, hydroxylates arenes, forms N-oxides, and converts sulfides into sulfones [Review: Muzart *Synthesis* 1325 1995], and converts selectively primary and secondary alcohols to their respective carbonyl compounds in the presence of catalytic amounts of K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> and Adogen 464 [see 72749-59-8] [Muzart et al. *Tetrahedron Lett* 35 1989 1994, Mohand & Muzart *Synth Commun* 25 2373 1995]. It is a useful disinfectant and antiseptic, and is best stored blow 10°. For X-ray crystallography, the crystals were prepared by slowly evaporating in air a solution of Na<sub>2</sub>CO<sub>3</sub> in 10% H<sub>2</sub>O<sub>2</sub>, or in an excess of 15% H<sub>2</sub>O<sub>2</sub> or D<sub>2</sub>O<sub>2</sub> by standing for 1 hour at -5° before precipitating with EtOH, washing with EtOH and Et<sub>2</sub>O, and drying in air. Na<sub>2</sub><sup>13</sup>C<sub>3</sub> was similarly used. The X-ray structure revealed the presence of H<sub>2</sub>O<sub>2</sub> molecules hydrogen bonded to the CO<sub>3</sub> ions in the crystal, with two types of H<sub>2</sub>O<sub>2</sub> molecules. The IR and Raman spectra of the normal, <sup>13</sup>C and <sup>2</sup>H isotopically substituted “percarbonates” are consistent with the X-ray data where the two H<sub>2</sub>O<sub>2</sub> sites are different. The IR (nujol, 25°) of Na<sub>2</sub>CO<sub>3</sub>·1.5H<sub>2</sub>O<sub>2</sub> has ν<sub>max</sub> at 3050w, 2900m, 2720sh, 2490s, 2350m, 1550m, 985m, 960s, 973vw, 868w cm<sup>-1</sup>, and Na<sub>2</sub>CO<sub>3</sub>·1.5<sup>2</sup>H<sub>2</sub>O<sub>2</sub> has ν<sub>max</sub> at 2470w, 2210m, 1940sh, 1890s, 1820m, 1130m, 1040s, 1035m, 873vw, 869w, 660m, 635s cm<sup>-1</sup> [Carrondo et al. *J Chem Soc, Dalton Trans* 2323 1977.]

**Sodium perchlorate (anhydrous)** [7601-89-0] **M 122.4, m 130°(for monohydrate),  $d_4^{20}$  2.02,  $pK_1^{25}$  -2.4 to -3.1 (for  $HClO_4$ ).** Because its solubility in water is high (2.1g/ml at 15°) and it has a rather low-temperature coefficient of solubility, sodium perchlorate is usually crystallised from acetone, MeOH, water/ethanol or dioxane/water (33g dissolved in 36ml of water and 200ml of dioxane is added). After filtering and recrystallising, the solid is dried under vacuum at 140-150° to remove solvent of crystallisation. Basic impurities can be removed by crystallisation from hot acetic acid, followed by heating at 150°. If  $NaClO_4$  is precipitated from distilled water by adding  $HClO_4$  to the chilled solution, the precipitate contains some free acid. **EXPLOSIVE.**

**Sodium phosphoamidate** [3076-34-4] **M 119.0.** Dissolve it in water below 10°, and acetic acid is added dropwise to pH 4.0 so that the monosodium salt is precipitated. The precipitate is washed with water and  $Et_2O$ , then dried in air. Addition of one equivalent of NaOH to the solution gives the sodium salt, the solution being adjusted to pH 6.0 before use [Rose & Heald *Biochem J* **81** 339 1961].

**Sodium-potassium alloy** [11135-81-2] **M depends on composition, liquid.** This is generally in the form of a liquid globule from which an aliquot can be pipetted under dry  $N_2$ . It is available commercially in two ratios, viz K (78wt%), Na (22 wt%) and K (56 wt%), Na (44 wt%), in ampoules. A fresh *ca* ~5:1  $NaK_{2.8}$  liquid alloy can be prepared from clean K metal (77g, 2.0mol) and clean Na (16g, 0.7mol) by heating gently in xylene (100ml, freshly distilled from Na) until the metals coalesce, then dry oxygen-free diglyme (50ml) is added in order to keep the alloy in one globule (stir with a glass rod). Small globules of the alloy can be united by adding a few drops of *iso*-PrOH to the mixture. Cool to room temperature, and this can be stored under dry  $N_2$  indefinitely. An aliquot can be safely pipetted out from the central globule by syringe equipped with a metal stopcock. The alloy can be more safely disposed of than the Na/Hg alloy, but must be handled carefully as it ignites in moist air. Residues are readily destroyed by a 1:1 mixture of *iso*-PrOH/petroleum ether (b 60-100°). [Ellis & Flom *J Organomet Chem* **99** 263 1975, Gilman & Young *J Org Chem* **1**, 315 1936.] When it reacts, e.g. in THF, to form metal carbonyl anions, and with some hydrocarbons, e.g. cyclopentadienes, diphenyl ether, it always forms the K-metallide, i.e. cyclopentadienylpotassium and  $C_6H_5K$ , and not the Na-metallide [Bryce-Smith & Turner *J Chem Soc* 861 1953, Ziegler & Schnell *Justus Liebigs Annalen Chemie* **437** 227 1924, Müller & Bunge *Chem Ber* **69** 2164 1936].

**Sodium pyrophosphate decahydrate** [13472-36-1] **M 446.1,  $d_4^{20}$  1.82,  $pK_1^{25}$  1.52,  $pK_2^{25}$  2.36,  $pK_3^{25}$  6.60,  $pK_4^{25}$  9.25 (for pyrophosphoric acid,  $H_4P_2O_7$ ).** Crystallise the salt from hot  $H_2O$  and dry it in air at room temperature.

**Sodium selenate** [13410-01-0] **M 188.9,  $d_4^{20}$  3.21,  $pK_1^{25}$  ~0,  $pK_2^{25}$  1.66 (for selenic acid,  $H_2SeO_4$ ).** Crystallise sodium selenate from hot water. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 433 1963].

**Sodium selenite** [10102-18-8] **M 172.9, m >350°,  $pK_1^{25}$  2.62,  $pK_2^{25}$  8.32 (for  $H_2SeO_3$ ).** Crystallise sodium selenite from a saturated aqueous solution where its solubility is 68% at 20° to give the *pentahydrate* salt. This yields the *anhydrous* salt on heating at 40°. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 431 1963].

**Sodium silicate solution** [1344-09-8]  **$pK_1^{25}$  9.51,  $pK_2^{25}$  11.77 (for silicic acid,  $H_4SiO_4$ )** Purify by contact filtration with activated charcoal.

**Sodium sulfate decahydrate** [7727-73-3 ( $10H_2O$ ), 7757-82-6 (*anhydrous*)] **M 322.2, m 32°(dec), 884° (anhydrous),  $d_4^{20}$  2.68 (anhydrous).** Crystallise sodium sulfate from water at 30° (1.1ml/g) by cooling to 0°. It becomes anhydrous at 32°.

**Sodium sulfide nonahydrate** [1313-84-4 ( $9H_2O$ ), 1313-82-2 (*anhydrous*)] **M 240.2, m ~50(loses  $H_2O$ ), 950(anhydrous),  $d_4^{20}$  1.43 ( $10H_2O$ ), 1.86 (anhydrous),  $pK_1^{25}$  7.04,  $pK_2^{25}$  11.96 (for  $H_2S$ ).** Some purification of the hydrated salt can be achieved by selecting large crystals and removing the surface layer (contaminated with oxidation products) by washing with distilled water. Other metal ions can be removed from  $Na_2S$  solutions

by passage through a column of Dowex ion-exchange A-1 resin, Na<sup>+</sup>-form. The hydrated salt can be rendered *anhydrous* by heating it in a stream of H<sub>2</sub> or N<sub>2</sub> until water is no longer evolved. (The resulting cake should not be heated to fusion because it is readily oxidised.) Recrystallise it from distilled water [Anderson & Azowlay *J Chem Soc, Dalton Trans* 469 1986]. Note that sodium sulfide hydrolyses in H<sub>2</sub>O to form NaHS + H<sub>2</sub>O, and is therefore alkaline. A 0.1N solution in H<sub>2</sub>O is 86% hydrolysed at room temperature. Its solubility in H<sub>2</sub>O is 8% at 0°, 12% at 20° and 30% at 50°. The *anhydrous* salt is obtained by allowing it to stand in a vacuum over conc H<sub>2</sub>SO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub> at 45° to start with, then at 30-35° when the salt contains 4% of water. The last traces of water are removed by heating to 700° in a glass or porcelain tube in a stream of H<sub>2</sub> to give pure H<sub>2</sub>S. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 358-360 1963.]

**Sodium sulfite** [7757-83-7] M 126.0, d<sub>4</sub><sup>20</sup> 2.63. Crystallise the sulfite from warm water (0.5ml/g) by cooling to 0°. Also purify it by repeated crystallisation from deoxygenated water inside a glove-box, and finally drying it under vacuum. [Rhee & Dasgupta *J Phys Chem* 89 1799 1985.]

**Sodium tetrametaphosphate** [13396-41-3] M 429.9, pK<sub>1</sub><sup>25</sup> 2.60, pK<sub>2</sub><sup>25</sup> 6.4, pK<sub>3</sub><sup>25</sup> 8.22, pK<sub>4</sub><sup>25</sup> 11.4 (tetrametaphosphoric acid, H<sub>4</sub>P<sub>4</sub>O<sub>12</sub>). Crystallise it twice from water at room temperature by adding EtOH (300g of Na<sub>4</sub>P<sub>4</sub>O<sub>12</sub>·H<sub>2</sub>O, 2L of water, and 1L of EtOH), wash it first with 20% EtOH then with 50% EtOH and dry it in air [Quimby *J Phys Chem* 58 603 1954].

**Sodium thioantimonate (Na<sub>3</sub>SbS<sub>4</sub>·9H<sub>2</sub>O, Schlippe's salt)** [13776-84-6] M 481.1, m 87°, b 234°, d<sub>4</sub><sup>20</sup> 1.81. Crystallise it from warm water (2ml/g made weakly alkaline with a few drops of dilute aqueous NaOH) by cooling to 0°. It forms a yellow *nonhydrate* which readily *effloresces* in air. [Schenk in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 619 1963.]

**Sodium thiocyanate** [540-72-7] M 81.1, m 300°, pK<sup>25</sup> -1.85 (for HSCN). It is recrystallised from EtOH or Me<sub>2</sub>CO, and the mother liquor is removed from the crystals by centrifugation. It is very *deliquescent* and should be kept in an oven at 130° before use. It can be dried in a vacuum at 120°/P<sub>2</sub>O<sub>5</sub> [Partington & Winterton *Trans Faraday Soc* 30 1104 1934]. Its solubility in H<sub>2</sub>O is 113% at 10°, 178% at 46°, 225.6% at 101.4°; in MeOH 35% at 15.8°, 51% at 48°, 53.5% at 52.3°; in EtOH 18.4% at 18.8°, 24.4% at 70.9°; and in Me<sub>2</sub>CO 6.85% at 18.8° and 21.4% at 56° [Hughes & Mead *J Chem Soc* 2282 1929].

Sodium thiocyanate has also been recrystallised from water, acetonitrile or from MeOH using Et<sub>2</sub>O for washing, then dried at 130°, or dried under vacuum at 60° for 2 days. [Strasser et al. *J Am Chem Soc* 107 789 1985, Szezygiel et al. *J Am Chem Soc* 91 1252 1987.] (The latter purification removes material reacting with iodine.) Sodium thiocyanate solutions can be freed from traces of iron by repeated batch extractions with Et<sub>2</sub>O.

**Sodium thiosulfate pentahydrate** [10102-17-7 (*hydrate*), 7772-98-7 (*anhydrous*)] M 248.2 (*anhydrous*), m 48 (*rapid heat*), d<sub>4</sub><sup>20</sup> 1.69, pK<sub>1</sub><sup>25</sup> 0.6, pK<sub>2</sub><sup>25</sup> 1.74 (for H<sub>2</sub>S<sub>2</sub>O<sub>3</sub>). Crystallise it from EtOH/H<sub>2</sub>O solutions or from water (0.3ml/g) below 60° by cooling to 0°, and dry it at 35° over P<sub>2</sub>O<sub>5</sub> under vacuum. [Foerster & Mommsen *Chem Ber* 57 258 1924.] This salt is used as a secondary standard in volumetric analysis [Kilpatrick *J Am Chem Soc* 45 2132 1923], and is used as "Hypo" in photography [Hargreaves & Dunningham *J Soc Chem Ind* 42 147T 1923.]

**Sodium trimetaphosphate hexahydrate** [7785-84-4] M 320.2, m 53°, d<sub>4</sub><sup>20</sup> 1.79, pK<sub>2</sub><sup>25</sup> 1.64, pK<sub>3</sub><sup>25</sup> 2.07 (for trimetaphosphoric acid, H<sub>3</sub>P<sub>3</sub>O<sub>9</sub>). It is precipitated from an aqueous solution at 40° by adding EtOH. It is dried in air.

**Sodium tripolyphosphate** [7758-29-4] M 367.9, pK<sub>1</sub><sup>25</sup> ~1, pK<sub>2</sub><sup>25</sup> 2.0, pK<sub>3</sub><sup>25</sup> 2.13, pK<sub>4</sub><sup>25</sup> 5.78, pK<sub>5</sub><sup>25</sup> 8.56 (for tripolyphosphoric acid, H<sub>5</sub>P<sub>3</sub>O<sub>10</sub>). Purify it by repeated precipitation from aqueous solution by slow addition of MeOH and dried in air. Also a solution of anhydrous sodium tripolyphosphate (840g) in water (3.8L) is filtered, MeOH (1.4L) is added with vigorous stirring to precipitate Na<sub>5</sub>P<sub>3</sub>O<sub>10</sub>·6H<sub>2</sub>O. The precipitate is collected on a filter, air dried by suction, then left to dry in air overnight. It is crystallised twice more in this way, using a 13% aqueous solution (w/w), and leaching the crystals with 200ml portions of water [Watters et al. *J Am Chem Soc* 78 4855 1956]. Similarly, EtOH can be added to precipitate the salt from a filtered 12-15% aqueous solution, the final solution containing *ca* 25% EtOH (v/v). Air drying should be at a relative humidity of 40-

60%. Heat and vacuum drying should be avoided. [Quimby *J Phys Chem* **58** 603 1954, Klement in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 547 1963.]

**Sodium tungstate dihydrate** [10213-10-2] **M 329.9, m 698°, d<sub>4</sub><sup>20</sup> 4.18, pK<sub>1</sub><sup>25</sup> 2.20, pK<sub>2</sub><sup>25</sup> 3.70 (for tungstic acid, H<sub>2</sub>WO<sub>4</sub>).** The salt crystallises from hot water (0.8ml/g) on cooling to 0°. [Grüttner & Jender in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1727 1965].

**Stannic chloride (tin IV chloride, stannic tetrachloride)** [7646-78-8] **M 260.5, m -33°, -30.2°, b 114°/760mm, d<sub>4</sub><sup>20</sup> 2.23, pK<sub>25</sub> 14.15 (for aquo Sn<sup>4+</sup> hydrolysis).** SnCl<sub>4</sub> fumes in moist air due to formation of a hydrate. Fractionate it in a ground glass still and store it in the absence of air. Possible impurities are SO<sub>2</sub> and HCl [Baudler in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 729 1963]. It forms a solid *pentahydrate* [10026-06-9] which smells of HCl and is obtained when the *anhydrous* salt is dissolved in a small volume of H<sub>2</sub>O. Also reflux it with clean mercury or P<sub>2</sub>O<sub>5</sub> for several hours, then distil it under (reduced) N<sub>2</sub> pressure into a receiver containing P<sub>2</sub>O<sub>5</sub>. Finally redistil it. *Alternatively*, distil it from Sn metal under vacuum in an all-glass system and seal off in large ampoules. SnCl<sub>4</sub> is available commercially as 1M solutions in CH<sub>2</sub>Cl<sub>2</sub> or hexane. **HARMFUL VAPOURS.**

**Stannic iodide (SnI<sub>4</sub>)** [7790-47-8] **M 626.3, m 144°, b 340, d<sub>4</sub><sup>20</sup> 4.46.** It is recrystallised from anhydrous CHCl<sub>3</sub>, dry it under vacuum and store it in a vacuum desiccator. It sublimes at 180°.

**Stannic oxide (SnO<sub>2</sub>)** [18282-10-5] **M 150.7, m 1630°, d<sub>4</sub><sup>20</sup> 6.95.** Reflux it repeatedly with fresh HCl until the acid shows no tinge of yellow. The oxide is then dried at 110°.

**Stannous chloride (anhydrous)** [7772-99-8] **M 189.6, m 247°, b 606°, d<sub>4</sub><sup>20</sup> 3.95, pK<sub>1</sub><sup>25</sup> 1.7, pK<sub>2</sub><sup>25</sup> 3.7 (for aquo Sn<sup>2+</sup> hydrolysis).** Analytical reagent grade stannous chloride *dihydrate* is dehydrated by adding it slowly to vigorously stirred, redistilled acetic anhydride (120g salt per 100g of anhydride) in a fume cupboard. After *ca* an hour, the anhydrous SnCl<sub>2</sub> is filtered on to a sintered-glass or Büchner funnel, washed free from acetic acid with dry Et<sub>2</sub>O (2 x 30ml), and dried under vacuum. It is stored in a sealed container. [Stephen *J Chem Soc* 2786 1930, Williams *Org Synth Coll Vol III* 627 1955.]

**Strontium bromide** [10476-81-0] **M 247.4, m 643°, d<sub>4</sub><sup>20</sup> 4.22.** Crystallise the bromide from water (0.5ml/g). It is *deliquescent*. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 930 1963.]

**Strontium chloride hexahydrate** [1025-70-4] **M 266.6, m 61°(rapid heating), 114-150°(loses 5H<sub>2</sub>O), 868°(anhydrous).** It crystallises from warm water (0.5ml/g) on cooling to 0°. It dehydrates at 150-160° in a vacuum. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 930 1963.]

**Strontium chromate** [7789-06-2] **M 203.6, d<sub>4</sub><sup>20</sup> 3.9, pK<sub>1</sub><sup>25</sup> 0.74, pK<sub>2</sub><sup>25</sup> 6.49 (for H<sub>2</sub>CrO<sub>4</sub>).** Crystallise strontium chromate from water (40ml/g) by cooling.

**Strontium hydroxide octahydrate** [1311-10-0 (8H<sub>2</sub>O), 18480-07-4 (anhydrous)] **M 265.8, m 100° (loses H<sub>2</sub>O), d<sub>4</sub><sup>20</sup> 1.90, m 375 (anhydrous), d 3.63 (anhydrous).** Crystallise the hydroxide from hot water (2.2ml/g) by cooling to 0°.

**Strontium nitrate** [10042-76-9] **M 211.6, m 570°, b 645°, d<sub>4</sub><sup>20</sup> 2.99.** Crystallise it from hot water (0.5ml/g) by cooling to 0°.

**Sulfamic acid** [5329-14-6] **M 97.1, m 205°(dec), pK<sub>25</sub> 0.99 (NH<sub>2</sub>SO<sub>3</sub>H).** Crystallise NH<sub>2</sub>SO<sub>3</sub>H from water at 70° (300ml per 25g), after filtering, by cooling a little and discarding the first batch of crystals (about 2.5g) before standing in an ice-salt mixture for 20 minutes. The crystals are filtered off by suction, washed with a small quantity of ice cold water, then twice with cold EtOH and finally with Et<sub>2</sub>O. Dry it in air for 1 hour, then

store it in a desiccator over  $\text{Mg}(\text{ClO}_4)_2$  [Butler et al. *Ind Eng Chem (Anal Ed)* **10** 690 1938]. For the preparation of primary standard material see *Pure Appl Chem* **25** 459 1969.

**Sulfamide** [7803-58-9] **M 96.1, m 91.5°**. Crystallise sulfamide from absolute EtOH. It decomposes at 250°. [Schenk in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** pp 482-483 1963.]

**Sulfur** [7704-34-9] **M 32.1, m between 112.8° and 120°, depending on form**. Murphy, Clabaugh & Gilchrist [*J Res Nat Bur Stand* **64A** 355 1960] have obtained sulfur of about 99.999% purity by the following procedure: Roll sulfur was melted and filtered through a coarse-porosity glass filter funnel into a 2L round-bottomed Pyrex flask with two necks. Conc  $\text{H}_2\text{SO}_4$  (300ml) was added to the sulfur (2.5kg), and the mixture was heated to 150°, stirring continuously for 2 hours. Over the next 6 hours, conc  $\text{HNO}_3$  was added in about 2ml portions at 10-15 minutes intervals to the heated mixture. It was then allowed to cool to room temperature and the acid was poured off. The sulfur was rinsed several times with distilled water, then remelted, cooled, and rinsed several times with distilled water again, this process being repeated four or five times to remove most of the acid entrapped in the sulfur. An air-cooled reflux tube (*ca* 40cm long) was attached to one of the necks of the flask, and a gas delivery tube (the lower end about 2.5cm above the bottom of the flask) was inserted into the other. While the sulfur was boiled under reflux, a stream of helium or  $\text{N}_2$  was passed through to remove any water,  $\text{HNO}_3$  or  $\text{H}_2\text{SO}_4$ , as vapours. After 4 hours, the sulfur was cooled so that the reflux tube could be replaced by a bent air-cooled condenser. The sulfur was then distilled, rejecting the first and the final 100ml portions, and transferred in 200ml portions to 400ml glass cylinder ampoules (which were placed on their sides during solidification). After adding about 80ml of water, displacing the air with  $\text{N}_2$ , the ampoule was cooled, and the water was titrated with 0.02M NaOH, the process being repeated until the acid content was negligible. Finally, entrapped water was removed by alternate evacuation to 10mm Hg and refilling with  $\text{N}_2$  while the sulfur was kept molten. The ampoules were then sealed. Other purifications include crystallisation from  $\text{CS}_2$  (which is less satisfactory because the sulfur retains appreciable amounts of organic material), \*benzene or \*benzene/acetone, followed by melting and degassing. It has also been boiled with 1% MgO, then decanted, and dried under a vacuum at 40° for 2days over  $\text{P}_2\text{O}_5$ . [For the purification of  $\text{S}_6$ , "recrystallised  $\text{S}_8$ " and "Bacon-Fanelli sulfur" see Bartlett et al. *J Am Chem Soc* **83** 103, 109 1961.]

**Sulfur dichloride (sulfur chloride,  $\text{SCl}_2$ )** [10545-99-0] **M 103.0, m -78°, b 59°/760mm(dec),  $d_4^{20}$  1.621**. Distil sulfur chloride twice in the presence of a small amount of  $\text{PCl}_3$  through a 12in Vigreux column, the fraction boiling between 55-61° being redistilled (in the presence of  $\text{PCl}_3$ ), and the fraction distilling between 58-61° retained. (The  $\text{PCl}_3$  is added to inhibit the decomposition of  $\text{SCl}_2$  into  $\text{S}_2\text{Cl}_2$  and  $\text{Cl}_2$ ). The  $\text{SCl}_2$  must be used as quickly as possible after distillation — within 1 hour at room temperature. The sample contains 4% of  $\text{S}_2\text{Cl}_2$ . On long standing this reaches 16-18%. [Fehér in *Handbook of Preparative Inorganic Chemistry (Ed. Brauer)* Academic Press Vol **I** pp 371-372 1963.] **HARMFUL VAPOURS.**

**Sulfur dioxide** [7446-09-5] **M 64.1, m -72°, b -10°**. Dry it by bubbling through conc  $\text{H}_2\text{SO}_4$  and by passage over  $\text{P}_2\text{O}_5$ , then through a glass-wool plug. Freeze it with liquid air and pump it to a high vacuum to remove dissolved gases. It is easily liquefied by compression (2.5 atmospheres at 15°), or by passing it through a glass spiral column in a freezing mixture of ice and salt. It is a colourless liquid with a density of 1.434 at 0°, which on rapid evaporation forms a snow white solid. It could be used as a solvent in certain reactions. **HARMFUL SUFFOCATING VAPOURS.**

**Sulfuric acid (oil of vitriol)** [7664-93-9] **M 98.08,  $d_4^{20}$  1.836 (96-98%), 1.805 (100%), m 3.0° (98%), 10.36° (100%), b 330.0 ± 0.5° (100%), dec > 340.0°, dihydrate M 134.11, m -38.9°, b 167°,  $\text{pK}_1^{25} \sim -8.3$ ,  $\text{pK}_2^{25} 1.99$** . Sulfuric acid, and also 30% fuming  $\text{H}_2\text{SO}_4$ , can be distilled in an all-Pyrex system, optionally from potassium persulfate. It has been purified by fractional crystallisation of the *monohydrate* from the liquid. It has a very strong dehydrating action and attacks skin—wash immediately with cold  $\text{H}_2\text{O}$  **and protect eyes**; otherwise the skin can be scarred for life **and could be blinding**. It is very *hygroscopic* and has been used as a desiccant in desiccators. Dilution with  $\text{H}_2\text{O}$  is highly exothermic, and because the concentrated acid is much more dense than  $\text{H}_2\text{O}$  it is diluted by running the concentrated acid down the side of the container of  $\text{H}_2\text{O}$  with slow stirring while cooling the outside of the container in an ice bath. If these precautions are not taken, the  $\text{H}_2\text{O}$  is likely to splatter and boil vigorously over. **CORROSIVE.**

**Sulfuric acid (fuming) (oleum)** [8014-95-7] **M 98.08 (+ x SO<sub>3</sub>)**. This is sulfuric acid with various amounts of SO<sub>3</sub> dissolved in it up to 80% free SO<sub>3</sub>. The specific gravities, **Sp gr**<sub>4</sub><sup>20</sup>, of oleums (% free SO<sub>3</sub>) are: 1.899 (15%), 1.915 (20%), 1.952 (30%), 2.001 (50%), 2.102 (60% maximum density) and 1.949 (80%). They are an almost colourless viscous liquids which emit choking fumes of SO<sub>3</sub>, and should be handled with great care: **use eye and body protection**. It reacts explosively with H<sub>2</sub>O to form strong sulfuric acid, and is a strong sulfonating agent.

**Sulfur monochloride (sulfur monochloride, S<sub>2</sub>Cl<sub>2</sub>)** [10025-67-9] **M 135.0, m -77°; b 19.1°, 29-30°/12mm, 72°/100mm, 138°/760mm, d**<sub>4</sub><sup>20</sup> **1.677, n**<sub>D</sub><sup>20</sup> **1.67**. It is a *pungent, irritating golden yellow liquid*. When impure its colour is orange to red due to SCl<sub>2</sub> formed. It fumes in moist air and liberates HCl, SO<sub>2</sub> and H<sub>2</sub>S in the presence of H<sub>2</sub>O. Distil it and collect the fraction boiling above 137° at atmospheric pressure. Fractionate this fraction over sulfur at *ca* 12mm using a ground glass apparatus (**b** 29-30°). *Alternatively*, purify it by distillation below 60° from a mixture containing sulfur (2%) and activated charcoal (1%), under reduced pressure (e.g. 50mm). It is soluble in EtOH, \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O, CS<sub>2</sub> and CCl<sub>4</sub>. Store it in a closed container in the dark in a refrigerator. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 371 1963.] **HARMFUL VAPOURS.**

**Sulfuryl chloride (SO<sub>2</sub>Cl<sub>2</sub>)** [7791-25-2] **M 135.0, m -54.1°, b 69.3°/760mm, d**<sub>4</sub><sup>20</sup> **1.67, n**<sub>D</sub><sup>20</sup> **1.44**. It is a *pungent, irritating colourless liquid*. It becomes yellow with time due to decomposition to SO<sub>2</sub> and HCl. Distil it and collect fraction boiling below 75°/atm which is mainly SO<sub>2</sub>Cl<sub>2</sub>. To remove HSO<sub>3</sub>Cl and H<sub>2</sub>SO<sub>4</sub> impurities, the distillate is poured into a separating funnel filled with crushed ice and *briefly* shaken. The lower cloudy layer is removed, dried for some time in a desiccator over P<sub>2</sub>O<sub>5</sub> and finally fractionate it at atmospheric pressure. The middle fraction boils at 69-70° and is pure SO<sub>2</sub>Cl<sub>2</sub>. It decomposes gradually in H<sub>2</sub>O to H<sub>2</sub>SO<sub>4</sub> and HCl. It reacts **violently** with EtOH and MeOH and is soluble in \*C<sub>6</sub>H<sub>6</sub>, toluene, Et<sub>2</sub>O and acetic acid. [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed Brauer) Academic Press Vol I p 383 1963, Allen & Maxson *Inorg Synth* I 114 1939]. **HARMFUL VAPOURS.**

**Tantalum (V) chloride (tantalum pentachloride)** [7721-01-9] **M 358.2, m 216.2°, 216.5-220°, b 239°/atm, d**<sub>4</sub><sup>20</sup> **3.68**. Purify it by sublimation in a stream of Cl<sub>2</sub>. It forms colourless needles when pure (yellow when contaminated with even less than 1% of NbCl<sub>5</sub>). It is sensitive to H<sub>2</sub>O; even in conc HCl it decomposes to tantalic acid. It is soluble in EtOH. [Rolsten *J Am Chem Soc* **80** 2952 1958, Brauer in *Handbook of Preparative Inorganic Chemistry* (Ed Brauer) Academic Press Vol II p 1302 1965.]

**Telluric acid** [7803-68-1] **M 229.6, pK**<sub>1</sub><sup>25</sup> **7.70, pK**<sub>2</sub><sup>25</sup> **11.04 (H<sub>6</sub>TaO<sub>6</sub>)**. Crystallise it once from nitric acid, then repeatedly from hot water (0.4ml/g). [Fehér in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I pp 451-453 1963.]

**Tellurium** [13494-80-9] **M 127.6, m 450°**. Purify tellurium by zone refining and repeated sublimation to an impurity of less than 1 part in 10<sup>8</sup> (except for surface contamination by TeO<sub>2</sub>). [Machol & Westrum *J Am Chem Soc* **80** 2950 1958.] Tellurium is volatile at 500°/0.2mm. It has also been purified by electrode deposition [Mathers & Turner *Trans Amer Electrochem Soc* **54** 293 1928].

**Tellurium dioxide** [7446-07-3] **M 159.6, m 733°, d**<sub>4</sub><sup>20</sup> **6.04**. Dissolve it in 5M NaOH, filter it and precipitate it by adding 10M HNO<sub>3</sub> (CARE) to the filtrate until the solution is acid to phenolphthalein. After decanting the supernatant, the precipitate is washed five times with distilled water, then dried for 24 hours at 110° [Horner & Leonhard *J Am Chem Soc* **74** 3694 1952].

**Terbium oxide** [12037-01-3] **M 747.7, pK**<sup>25</sup> **8.16 (for Tb<sup>3+</sup> hydrolysis)**. Dissolve it in acid (e.g. perchloric acid), precipitate it as its oxalate and ignite the oxalate at 650°.

**Thallium (III) nitrate trihydrate** [TTN, thallic nitrate trihydrate,  $\text{Tl}(\text{NO}_3)_3 \cdot 3\text{H}_2\text{O}$ ] [13453.38-8] **M 444.4, m 102-105°**. It is prepared from Tl(III)oxide (45g, [1314-32-5]) by stirring it in hot concentrated  $\text{HNO}_3$  (120ml) at  $\sim 50^\circ$  for 30 minutes when all the oxide dissolves to give a clear colourless solution. On cooling to  $0^\circ$ , a white solid crystallises out. It is filtered off, washed with a little cold  $\text{H}_2\text{O}$ , dilute  $\text{HNO}_3$ , dried in a vacuum desiccator over  $\text{P}_2\text{O}_5$  to give TTN (75g, 85%) as hard, colourless crystals which should be stored in a cold tightly sealed container. Treat carefully as it is **POISONOUS**. [McKillop et al. *J Am Chem Soc* **95** 3635 1973.]

**Thallos bromide** [7789-40-4] **M 284.3, m 460°**. Thallos bromide (20g) is purified by refluxing for 2-3 hours with water (200ml) containing 3ml of 47% HBr. It is then washed until acid-free, heated to  $300^\circ$  for 2-3 hours and stored in brown bottles. Its solubility in  $\text{H}_2\text{O}$  (w/w) is 0.034% at  $0^\circ$ , 0.048% at  $20^\circ$ , and 0.204% at  $100^\circ$ . [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 870 1963]. **POISONOUS**.

**Thallos carbonate** [6533-73-9] **M 468.7, m 268-270°**. It crystallises from hot water (4ml/g) on cooling. **POISONOUS**.

**Thallos chlorate** [13453-30-0] **M 287.8,  $d_4^{20}$  5.05**. It crystallises from hot water (2ml/g) on cooling. Its solubility in  $\text{H}_2\text{O}$  (w/w) is 0.17% at  $0^\circ$ , 0.32% at  $20^\circ$ , and 2.4% at  $100^\circ$ . [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 870 1963]. **POISONOUS**.

**Thallos chloride** [7791-12-0] **M 239.8, m 429.9°, b 806°,  $d_4^{20}$  7.0**. Recrystallise it from 1% HCl and wash it until acid-free, or crystallise it from hot water (50ml/g), then dry it at  $140^\circ$  and store it in brown bottles. Also purify it by subliming it in a vacuum, followed by treatment with dry HCl gas and filtering while molten. (It is soluble in 260 parts of cold water and 70 parts of boiling water). **POISONOUS**.

**Thallos hydroxide** [12026-06-1] **M 221.4, m 139°(dec),  $\text{pK}^{25}$  13.2 (for  $\text{Tl}^+$ )**. It crystallises from hot water (0.6ml/g) on cooling. **POISONOUS**.

**Thallos iodide** [7790-30-9] **M 331.3, m 441.8°, b 824°,  $d_4^{20}$  7.1**. Reflux it for 2-3 hours with water containing HI, then wash it until acid-free, and dry it at  $120^\circ$ . Store it in brown bottles. [Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 870 1963.] **POISONOUS**.

**Thallos nitrate** [10102-45-1] **M 266.4, m 206°, b 450°(dec),  $d_4^{20}$  5.55**. The nitrate crystallises from warm water (1ml/g) on cooling to  $0^\circ$ . **POISONOUS**.

**Thallos perchlorate** [13453-40-2] **M 303.8,  $\text{pK}^{25}$  -2.4 to -3.1 (for  $\text{HClO}_4$ )**. It crystallises from hot water (0.6ml/g) on cooling. Dry it under vacuum for 12 hours at  $100^\circ$  (protect from possible **EXPLOSION**).

**Thallos sulfate** [7446-18-6] **M 504.8, m 633°,  $d_4^{20}$  6.77**. The sulfate crystallises from hot water (7ml/g) by cooling; then dry it under vacuum over  $\text{P}_2\text{O}_5$ . It is **POISONOUS**.

**Thionyl chloride** [7719-09-7] **M 119.0, b 77°,  $d_4^{20}$  1.636**. Crude  $\text{SOCl}_2$  can be freed from sulfuryl chloride, sulfur monochloride and sulfur dichloride by refluxing it with sulfur and then fractionally distilling twice. [The  $\text{SOCl}_2$  is converted to  $\text{SO}_2$  and sulfur chlorides. The  $\text{S}_2\text{Cl}_2$  (**b** 135.6°) is left in the residue, whereas  $\text{SCl}_2$  (**b** 59°) passes over in the forerun.] The usual purification is to distil it from quinoline (50g  $\text{SOCl}_2$  to 10g quinoline) to remove acid impurities, followed by distillation from boiled linseed oil (50g  $\text{SOCl}_2$  to 20g of oil). Precautions must be taken to exclude moisture.

Thionyl chloride is used extensively in organic syntheses and can be prepared by distillation of technical  $\text{SOCl}_2$  in the presence of diterpene (12g/250ml  $\text{SOCl}_2$ ), and avoiding overheating. Further purification is achieved by redistillation from linseed oil (1-2%) [Rigby *Chem Ind (London)* 1508 1969]. Gas chromatographically pure material is obtained by distillation from 10% (w/w) triphenyl phosphite [Friedman & Wetter *J Chem Soc (A)* 36 1967, Larsen et al. *J Am Chem Soc* **108** 6950 1986]. **HARMFUL VAPOURS**.

**Thorium chloride** [10026-08-1] **M 373.8**,  $\text{pK}_1^{25}$  10.45,  $\text{pK}_2^{25}$  10.62,  $\text{pK}_3^{25}$  10.80,  $\text{pK}_4^{25}$  11.64 (for aquo  $\text{Th}^{4+}$ ). It is freed from anionic impurities by passing a 2M solution of  $\text{ThCl}_4$  in 3M HCl through a Dowex-1 anion-resin column. The eluate is partially evaporated to give crystals which are filtered off, washed with  $\text{Et}_2\text{O}$  and stored in a desiccator over  $\text{H}_2\text{SO}_4$  to dry. *Alternatively*, a saturated solution of  $\text{ThCl}_4$  in 6M HCl is filtered through quartz wool and extracted twice with ethyl, or isopropyl ether (to remove iron), then evaporated to a small volume on a hot plate. (Excess silica precipitates and is filtered off. The filtrate is cooled to  $0^\circ$  and saturated with dry HCl gas.) It is shaken with an equal volume of  $\text{Et}_2\text{O}$ , shaken with HCl gas, until the mixture becomes homogeneous. On standing,  $\text{ThCl}_4 \cdot 8\text{H}_2\text{O}$  precipitates out and is filtered off, washed with  $\text{Et}_2\text{O}$  and dried [Kremer *J Am Chem Soc* **64** 1009 1942].

**Thorium sulfate tetrahydrate** [10381-37-0] **M 496.2**, **m  $42^\circ$** (loses  $\text{H}_2\text{O}$ ),  $\text{d}_4^{20}$  2.8. Crystallise it from water. The solubility of the *decahydrate* increases with increase in temperature, whereas the solubility of the *tetrahydrate* decreases with increase of temperature.

**Tin** (powder) [7440-31-5] **M 118.7**. Tin powder is purified by adding it to about twice its weight of 10% aqueous NaOH and shaking vigorously for 10 minutes. (This removes oxide film and stearic acid or similar material that is sometimes added for pulverisation.) It is then filtered, washed with water until the washings are no longer alkaline to litmus, rinsed with MeOH and dried in air. [Sisido et al. *J Am Chem Soc* **83** 538 1961.]

**Titanium tetrabromide** [7789-68-6] **M 367.5**, **m  $28.3^\circ$** , **b  $233.5^\circ$** ,  $\text{d}_4^{20}$  3.3. Purify it by distillation. The distillate forms light orange hygroscopic crystals. Store it in the dark under  $\text{N}_2$  preferably in sealed brown glass ampules. [Olsen & Ryan *J Am Chem Soc* **54** 2215 1932.]

**Titanium tetrachloride** [7550-45-0] **M 189.7**, **b  $136.4^\circ$** ,  **$154^\circ$** , **d 1.730**,  $\text{pK}_1^{25}$  0.3,  $\text{pK}_2^{25}$  1.8,  $\text{pK}_3^{25}$  2.1,  $\text{pK}_4^{25}$  2.4 (for aquo  $\text{Ti}^{4+}$  hydrolysis). Reflux it with mercury or a small amount of pure copper turnings to remove the last traces of colour [due to  $\text{FeCl}_3$  and  $\text{VCl}_4$ ], then distil it under  $\text{N}_2$  in an all-glass system, taking precautions to exclude moisture. Clabaugh et al. [*J Res Nat Bur Stand* **55** 261 1955] removed organic material by adding aluminium chloride hexahydrate as a slurry with an equal amount of water (the slurry being *ca* one-fiftieth the weight of  $\text{TiCl}_4$ ), refluxed it for 2-6 hours while bubbling in chlorine, the excess of which is subsequently removed by passing a stream of clean dry air. The  $\text{TiCl}_4$  is then distilled, refluxed with copper and again distilled, taking precautions to exclude moisture. Volatile impurities are then removed using a technique of freezing, pumping and melting. The *titanium tetrachloride 2-tetrahydrofuran complex*,  $\text{TiCl}_4(\text{THF})_2$ , [31011-57-1] **M 333.9**, has **m  $126-128^\circ$**  and is easier to handle than  $\text{TiCl}_4$  [Hamilton et al. *Organometallics* **75** 2881 1953, Abrahamson et al. *Organometallics* **3** 1379 1984, Beilstein **17/1** V 33]. [Baxter & Fertig *J Am Chem Soc* **45** 1228 1923, Baxter & Butler *J Am Chem Soc* **48** 3117 1926.] **HARMFUL VAPOURS.**

**Titanium trichloride** [7705-07-9] **M 154.3**, **m  $>500^\circ$** ,  $\text{pK}_1^{25}$  2.55 (for hydrolysis of  $\text{Ti}^{3+}$  to  $\text{TiOH}^{2+}$ ). It is a brown purple powder that is very reactive to  $\text{H}_2\text{O}$  and pyrophoric when dry. It should be manipulated in a dry box. It is soluble in  $\text{CH}_2\text{Cl}_2$  and tetrahydrofuran, and is used as a M solution in these solvents in the ratio of 2:1, and stored under  $\text{N}_2$ . It is a powerful reducing agent. [Ingraham et al. *Inorg Synth* **VI** 52 1960.]

**Titanyl sulfate ( $\text{TiOSO}_4 \cdot 2\text{H}_2\text{O}$ )** [13825-74-6] **M 160.0**. Dissolve it in water, filter and crystallise it three times from boiling 45%  $\text{H}_2\text{SO}_4$ , washing with EtOH to remove excess acid, then with  $\text{Et}_2\text{O}$ . Dry it in air for several hours, then in an oven at  $105-110^\circ$ . [Hixson & Fredrickson *Ind Eng Chem* **37** 678 1945.]

**Triiron dodecacarbonyl** [17685-52-8] **M 503.7**, **m  $140^\circ$** (dec). It usually contains 10% by weight of MeOH as stabiliser. This can be removed by keeping it in a vacuum at 0.5mm for at least 5 hours. It can be sublimed slowly at high vacuum and is soluble in organic solvents. [Landesberg et al. *J Org Chem* **37** 930 1972, Case & Whiting *J Chem Soc* 4632 1960, King & Stone *Inorg Synth* **VII** 193 1963.] **TOXIC.**

**Trisodium orthophosphate dodecahydrate** [10101-89-0] **M 380.1**,  $\text{pK}_1^{25}$  2.15,  $\text{pK}_2^{25}$  7.21,  $\text{pK}_3^{25}$  12.33 (for  $\text{H}_3\text{PO}_4$ ). It crystallises from warm dilute aqueous NaOH (1ml/g) on cooling to  $0^\circ$ .

**Tritium** [10028-17-8] **M 6.0**. Purify tritium from hydrocarbons and  $^3\text{H}$  by diffusion through the wall of a hot nickel tube [Landecker & Gray *Rev Sci Instrum* **25** 1151 1954]. **RADIOACTIVE**.

**Tungsten** (rod) [7440-33-7] **M 183.6, m 3410°, b 5900°, d<sub>4</sub><sup>20</sup> 19.0**. Clean the solid with conc NaOH solution, rub it with very fine emery paper until its surface is bright, wash it with previously boiled and cooled conductivity water and dry it with filter paper. [Hein & Herzog in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1417 1965.]

**Tungsten hexacarbonyl** [14040-11-0] **M 351.9, d<sub>4</sub><sup>20</sup> 2.650**. Sublime it *in vacuo* before use [Connoe et al. *J Chem Soc, Dalton Trans* 511 1986]. **TOXIC**.

**Tungsten (VI) trichloride** [13283-01-7] **M 396.6, m 265°(dec), 275°, b 346°, d<sub>4</sub><sup>25</sup> 3.520, pK<sub>1</sub><sup>25</sup> 2.20, pK<sub>2</sub><sup>25</sup> 3.70 (for tungstic acid, H<sub>2</sub>WO<sub>4</sub>)**. Sublime it in a stream of Cl<sub>2</sub> in a high temperature furnace and collect it in a receiver cooled in a Dry-ice/acetone bath in an inert atmosphere because it is sensitive to moisture. It is soluble in CS<sub>2</sub>, CCl<sub>4</sub>, CHCl<sub>3</sub>, POCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub>, petroleum ether and Me<sub>2</sub>CO. Its solutions decompose on standing. Good crystals can be obtained by heating WCl<sub>6</sub> in CCl<sub>4</sub> to 100° in a sealed tube, followed by slow cooling (tablets of four-sided prisms). Store it in a desiccator over H<sub>2</sub>SO<sub>4</sub> in the dark. [Leitzke & Holt *Inorg Synth* **III** 163 1950, Parterfield & Tyree *Inorg Synth* **IX** 133 1967, Hein & Herzog in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1420 1965.]

**Uranium hexafluoride** [7783-81-5] **M 352.0, b 0°/17.4mm, 56.2°/765mm, m 64.8°, pK<sup>25</sup> 1.68 (for hydrolysis of U<sup>4+</sup> to UOH<sup>3+</sup>)**. Purify uranium hexafluoride by fractional distillation to remove HF. Also purify it by low-temperature trap-to-trap distillation over pre-dried NaF [Anderson & Winfield *J Chem Soc, Dalton Trans* 337 1986].

**Uranium trioxide** [1344-58-7] **M 286.0, d<sub>4</sub><sup>20</sup> 7.29**. The oxide is dissolved in HClO<sub>4</sub> (to give a uranium content of 5%), and the solution is adjusted to pH 2 by addition of dilute ammonia. Dropwise addition of 30% H<sub>2</sub>O<sub>2</sub>, with rapid stirring, precipitated U(VI) peroxide, the pH being held constant during the precipitation, by addition of small amounts of the ammonia solution. Then H<sub>2</sub>O<sub>2</sub> is added until further quantities caused no change in pH. After stirring for 1 hour, the slurry is filtered through coarse filter paper in a Büchner funnel, washed with 1% H<sub>2</sub>O<sub>2</sub> acidified to pH 2 with HClO<sub>4</sub>, then heated at 350° for three days in a large platinum dish [Baes *J Phys Chem* **60** 878 1956].

**Uranyl nitrate (UO<sub>2</sub>(NO<sub>3</sub>)<sub>2</sub> 6H<sub>2</sub>O)** [13520-83-7] **M 502.1, m 60.2°, b 118°, d<sub>4</sub><sup>25</sup> 2.807, pK<sup>25</sup> 5.82 (for aquo UO<sub>2</sub><sup>2+</sup>)**. Crystallise the nitrate from water by cooling to -5°, taking only the middle fraction of the solid which separates. Dry the *deliquescent* rhombic yellow crystals of the *hexahydrate* over 35-40% H<sub>2</sub>SO<sub>4</sub> in a vacuum desiccator. The crystals reflect a greenish lustre. They are remarkable because on crushing, rubbing or shaking they show *triboluminescence* with occasional detonation. They are very soluble in EtOH, and solutions of the nitrate in Et<sub>2</sub>O can explode in the presence of sunlight.

**Vanadium (metal)** [7440-62-2] **M 50.9, m 1910°, d<sub>4</sub><sup>20</sup> 6.0**. Clean the metal by rapid exposure consecutively to HNO<sub>3</sub>, HCl, HF, de-ionised water and reagent grade acetone, then dry it in a vacuum desiccator. [Brauer in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** pp 1252-1255 1965.]

**Vanadyl trichloride (VOCl<sub>3</sub>)** [7727-18-6] **M 173.3, m-79.5°, b 124.5-125.5°/744mm, 127.16°/760mm, d<sup>0</sup> 1.854, d<sup>32</sup> 1.811**. VOCl<sub>3</sub> should be lemon yellow in colour. If it is red, it may contain VCl<sub>4</sub> and Cl<sub>2</sub>. Fractionally distil it, and then redistil it over metallic Na, but be careful to leave some residue because the residue can become **EXPLOSIVE** in the presence of the metal. **USE A SAFETY SHIELD** and avoid contact with moisture. It readily hydrolyses to vanadic acid and HCl. Store it in a tightly closed container or in sealed ampoules under N<sub>2</sub>. [Brown & Griffiths *Inorg Synth* **I** 106 1939, Brown & Griffiths *Inorg Synth* **IV** 80 1953.]

**Water** [7732-18-5] **M 18.0, m 0<sup>0</sup>, b 100<sup>0</sup>, pK<sup>25</sup> 14.00.** Conductivity water (specific conductance *ca* 10<sup>-7</sup> mho) can be obtained by distilling water in a steam-heated tin-lined still, then, after adding 0.25% of solid NaOH and 0.05% of KMnO<sub>4</sub>, distilling once more from an electrically heated Barnstead-type still, taking the middle fraction into a Jena glass bottle. During these operations suitable traps must be used to protect against entry of CO<sub>2</sub> and NH<sub>3</sub>. Water, only a little less satisfactory for conductivity measurements (but containing traces of organic material) can be obtained by passing ordinary distilled water through a mixed bed ion-exchange column containing, for example, Amberlite resins IR 120 (cation exchange) and IRA 400 (anion exchange), or Amberlite MB-1. This treatment is also a convenient one for removing traces of heavy metals. (The metals Cu, Zn, Pb, Cd and Hg can be tested for by adding pure concentrated ammonia to 10ml of sample and shaking vigorously with 1.2ml of 0.001% dithizone in CCl<sub>4</sub>. Less than 0.1µg of metal ion will impart a faint colour to the CCl<sub>4</sub> layer.) For almost all laboratory purposes, simple distillation yields water of adequate purity, and most of the volatile contaminants such as ammonia and CO<sub>2</sub> are removed if the first fraction of distillate is discarded. Most laboratories have glass stills that “doubly” or “triple” distil water. [See “water” in Chapter 1.]

**Wood's Metal** [76093-98-6] **M ~1486, m ~75<sup>0</sup>, 73-77<sup>0</sup>.** It is a fusible alloy of Bi, Pb, Sn, Cd with the atomic composition of 4:2:1:1 (**m 71<sup>0</sup>**). It is commonly used as a metal bath for heating reaction flasks as it melts at a convenient temperature. Its temperature rises rather rapidly when heated with a Bunsen burner in a crucible (e.g. steel or nickel), so care should be taken when heating a reaction or distilling flask containing chemical reactants immersed into it. Heating should be applied slowly, and the temperature should be monitored continuously with a thermometer immersed in the molten metal. The temperature of the metal can be taken up to ~350<sup>0</sup>, and it is better to work in an efficient fume cupboard as the vapours are toxic. At higher temperature some oxidation of the metal takes place. A bath that has been heated at high temperatures and then the metal cooled back to the solid, and melted again has a solid scum of the oxides on its surface. These can be easily removed with a nickel spatula to leave a nice clean surface. The melting point of the re-used alloy is not seriously affected, and the alloy can be used over and over again without considerable loss in weight. It is commercially available in a granular form or as sticks with the composition of Bi(50 wt%), Pb (25 wt%), Sn (12.5 wt%) and Cd (12.5 wt%). Other such alloys are **Rose's metal (m 93.75<sup>0</sup>)** with the composition of Bi, Pb, Sn in the ratio of 2:1:1, and **Lipowitz alloy (m 60-65<sup>0</sup>)** with the composition of Bi, Pb, Sn, Cd in the ratio 15:8:4:3, but they are not commonly used in the laboratory. Alloys of Bi, Pb and Sn which melt slightly above 100<sup>0</sup> have been used as melting plugs in automatic sprinkler fire alarms.

**Zinc (dust)** [7440-66-6] **M 65.4.** Commercial zinc dust (1.2kg) is stirred with 2% HCl (3L) for 1 minute, then the acid is removed by filtration, and washed in a 4L beaker with a 3L portion of 2% HCl, three 1L portions of distilled water, two 2L portions of 95% EtOH, and finally with 2L of absolute Et<sub>2</sub>O. (The wash solutions were removed each time by filtration.) The material is then dried thoroughly, and if necessary, any lumps are broken up in a mortar. [Wagenknecht & Juza *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1067 1965.]

**Zinc (metal)** [7440-66-6] **M 65.4, m 420<sup>0</sup>, d<sub>4</sub><sup>20</sup> 7.141.** Fuse it under vacuum, cool it, then wash it with acid to remove the oxide.

**Zinc bromide** [7699-45-8] **M 225.2, m 384, b 697.** Heat ZnBr<sub>2</sub> to 300<sup>0</sup> under vacuum (2x10<sup>-2</sup> mm) for 1 hour, then sublime it. [Wagenknecht & Juza *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1072 1965.]

**Zinc chloride** [7646-85-7] **M 136.3, m 283<sup>0</sup>, 290<sup>0</sup>.** The anhydrous material can be sublimed under a stream of dry HCl, followed by heating to 400<sup>0</sup> in a stream of dry N<sub>2</sub>. It sublimes at high vacuum. Also purify it by refluxing (50g) in dioxane (400ml) with 5g zinc dust, filtering hot and cooling to precipitate ZnCl<sub>2</sub>. Crystallise it from dioxane and store it in a desiccator over P<sub>2</sub>O<sub>5</sub>. It has also been dried by refluxing in thionyl chloride. [Weberg et al. *J Am Chem Soc* **108** 6242 1986.] *Hygroscopic: minimal exposure to the atmosphere is necessary.* [Wagenknecht & Juza *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol II p 1070 1965.]

**Zinc-Copper Couple** [53801-63-1] **m 419-420°**. It is available in the form of a powder made of Zn and Cu. If the Zn contains some ZnO, then it should be treated with dilute HCl for a very short period, washed H<sub>2</sub>O, a little EtOH then Et<sub>2</sub>O, and dried *in vacuo* (see also zinc dust above). The *couple* can be prepared in either of the following ways. (1) Zinc dust (120g, MW of Zn is 65.4) and powdered copper oxide (10g, MW of Cu is 63.6) in a round-bottomed flask (~200ml) are heated gently over a flame in a current of H<sub>2</sub> and N<sub>2</sub> with stirring or rotating the flask until the CuO is reduced, and a uniformly gray mixture is obtained. The temperature should be maintained below the point of fusion during heating (~500°) and the mixture is stored in a desiccator. (2) Zinc powder (24.6g) is stirred with 3% aqueous HCl (20ml) for 1 minute, washed by decantation with 3% aqueous HCl (3 x 20ml), with H<sub>2</sub>O (5 x 20ml), with 2% aqueous CuSO<sub>4</sub> (2 x 40ml), distilled H<sub>2</sub>O (5 x 20ml), with EtOH (4 x 20ml), with Et<sub>2</sub>O (2 x 20ml), filtered off on a Büchner funnel, sucked dry then stored in a desiccator over P<sub>2</sub>O<sub>5</sub> overnight. [Corbin et al. *Org Synth* **44** 30 1964, Coll Vol **V** 328 1973; *Smith & Simmons Org Synth* **41** 72 1961, Coll Vol **V** 855 1973.] (3) A suitable zinc/copper alloy with 5—8% of Cu can be prepared by melting zinc (m 419.5°) with clean brass turnings (*ca* 2Cu/1Zn) and casting into bars, which can be turned into fine shavings prior to use. [Noller *Org Synth Coll Vol II* 185 1943.] Note that the couple loses activity in moist air, so it should be kept dry, and used as soon as possible after it is prepared for best results.

When it is required to reduce organic halides to the respective hydrocarbon or deuterated derivative, it is best to prepare the Zn-Cu couple in an oxygen-free system. Thus Zn dust (6.5g, 100mmol) suspended in H<sub>2</sub>O (10ml) is vigorously stirred with acidic cupric chloride solution (15ml of 0.15M in 5% hydrochloric acid) until evolution of gas ceases, the black solid is filtered off, and washed with H<sub>2</sub>O until the filtrate gives a negative chloride test with 6% AgNO<sub>3</sub> solution. The *couple* is washed twice with Me<sub>2</sub>CO, and in order to obtain the highest deuterium incorporation, the Me<sub>2</sub>CO wash should be followed by a D<sub>2</sub>O wash, two Me<sub>2</sub>CO washes, two Et<sub>2</sub>O washes and dried *in vacuo* at ~25°. This preparation of the *couple* is exceptional for replacing the halogen in a variety of organic halides by H or D in water-containing ether solvents, under mild conditions, and in reproducibly moderate to high yields (e.g. 1,4-dideuterobutadiene from 1,4-dichloro-1,3-butadiene, or 2-butanone-3-*d*<sub>1</sub> from 3-bromo-2-butanone). The deuterium derivatives provide a convenient <sup>13</sup>C NMR method for deuterium analysis. [Stephenson et al. *J Org Chem* **42** 212 1977.]

**Zinc cyanide** [557-21-1] **M 117.4, m 800°(dec), d<sub>4</sub><sup>20</sup> 1.852**. It is a **POISONOUS** white powder which becomes black on standing if Mg(OH)<sub>2</sub> and carbonate are not removed in the preparation. Thus, wash it well with H<sub>2</sub>O, then well with EtOH, Et<sub>2</sub>O and dry it in air at 50°. Analyse it by titrating the cyanide with standard AgNO<sub>3</sub>. Other likely impurities are ZnCl<sub>2</sub>, MgCl<sub>2</sub> and traces of basic zinc cyanide; the first two salts can be washed out. It is soluble in aqueous KCN solutions. However, if purified in this way Zn(CN)<sub>2</sub> is not reactive in the Gattermann synthesis. For this, the salt should contain at least 0.33 mols of KCl or NaCl which will allow the reaction to proceed faster. [Adams & Levine *J Am Chem Soc* **45** 2375 1923, Arnold & Sorung *J Am Chem Soc* **60** 1699 1938, Fuson et al. *Org Synth Coll Vol III* 549 1955.]

**Zinc fluoride** [7783-49-5] **M 103.4, m 872°, b 1500°, d<sub>4</sub><sup>25</sup> 5.00**. A possible impurity is H<sub>2</sub>O which can be removed by heating at 100° or by heating to 800° in a dry atmosphere. Heating in the presence of NH<sub>4</sub>F produces larger crystals. It is sparingly soluble in H<sub>2</sub>O (1.51g/100ml) but more soluble in HCl, HNO<sub>3</sub> and NH<sub>4</sub>OH. It can be stored in glass bottles. [Kwasnik in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 242 1963.]

**Zinc iodide** [10139-47-6] **M 319.2, m 446, b 624°(dec), d<sub>4</sub><sup>20</sup> 4.74**. Heat the iodide to 300°/2 x 10<sup>-2</sup>mm for 1 hour, then sublime it. Its solubility in H<sub>2</sub>O is 0.3ml/g, and it is soluble in EtOH. Store it in the dark.

**Zinc perchlorate hexahydrate** [13637-61-1] **M 372.4, m 105-107°, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>)**. Crystallise it from a small volume of H<sub>2</sub>O. It is soluble in EtOH. **Potentially EXPLOSIVE.**

**Zinc-Silver Couple** [37350-66-6]. It is best to prepare the couple fresh when used to replace a vinylic bromine atom in organic compounds by a hydrogen atom, deuterium atom or a tritium atom. The zinc dust to be used should first be cleaned from any zinc oxide or greasy material. Zinc dust (22g) is stirred under N<sub>2</sub> at ~25° as a suspension in 10% aqueous HCl (110ml) for 20 minutes. Acetone (50ml) is added to the mixture, the liquid is decanted from the metal, and the gray slurry is washed thoroughly with Me<sub>2</sub>CO (6 x 50ml) by decantation. The residual Zn is treated with a boiling suspension of AgOAc (~6.6g) in AcOH (125ml), stirred rapidly without

further heating, the liquid is decanted off and the black granular solid that is formed is washed with AcOH (2 x 60ml), Me<sub>2</sub>CO (2 x 20ml), and finally dry THF (2 x 50ml). Store it under N<sub>2</sub> in the dark. If it is used for deuterium replacement of bromine it is best to stir the couple with D<sub>2</sub>O (5ml) in dry THF (50ml) for 15 minutes, decant, repeat the process a second time and then wash the solid with dry THF (50ml). This reduces the amount of proton incorporation while maximising deuterium incorporation. [Fryzuk & Bosnich *J Am Chem Soc* **101** 3043 1979, Clark & Heathcock *J Org Chem* **41** 636 1976.]

**Zinc sulfate heptahydrate (white vitriol)** [7446-20-0 (7 H<sub>2</sub>O) 7446-19-7 (H<sub>2</sub>O), 7733-02-0 (anhydrous)] **M 287.5, m 100°(dec), 280°(loses all 7H<sub>2</sub>O), >500(anhydrous), d<sub>4</sub><sup>20</sup> 1.97.** Crystallise it from aqueous EtOH or dilute H<sub>2</sub>SO<sub>4</sub> below 39° when it forms the *heptahydrate*, and between 39° and 70° it forms the *hexahydrate*, and above 70° the *monohydrate* is stable. The *anhydrous salt* is obtained from the hydrates by heating at 280° or lower temperatures in a current of dry air. It decomposes to ZnO and SO<sub>2</sub> at 767°. The solubility of the *heptahydrate* in H<sub>2</sub>O is 5.88% at 0°, 61.92% at 30°, 66.61% at 35° and 70.05% at 39°.

**Zirconium tetrachloride** [10026-11-6] **M 233.0, m 300°(sublimes), pK<sub>1</sub><sup>25</sup> -0.32, pK<sub>2</sub><sup>25</sup> 0.06, pK<sub>3</sub><sup>25</sup> 0.35, pK<sub>4</sub><sup>25</sup> 0.46 (for hydrolysis of aquo Zr<sup>4+</sup>).** Crystallise it repeatedly from conc HCl. It is hydrolysed by H<sub>2</sub>O to form white ZrOCl<sub>2</sub> (see below) and HCl. [Krebs *Z Anorg Allgem Chem* **378** 263 1970.]

**Zirconyl chloride hexahydrate** [7699-43-6] **M 286.2, m 150°(loses 6H<sub>2</sub>O).** Crystallise it repeatedly from 8M HCl to give ZrOCl<sub>2</sub>·8H<sub>2</sub>O (see below). On drying, ZrOCl<sub>2</sub>·6H<sub>2</sub>O, **m 150°**, is formed. The product is not free from hafnium. [Blumenthal *J Chem Ed* **39** 607 1962.]

**Zirconyl chloride octahydrate** [13520-92-8] **M 322.3, m 150° (loses 6H<sub>2</sub>O), 210° (loses all H<sub>2</sub>O), 400° (anhydrous dec), d<sub>4</sub><sup>20</sup> 1.91.** Recrystallise the chloride several times from water [Ferragina et al. *J Chem Soc, Dalton Trans* 265 1986]. Recrystallisation from 8M HCl gives the *octahydrate* as white needles on concentrating. It is also formed by hydrolysing ZrCl<sub>4</sub> with water. After one recrystallisation from H<sub>2</sub>O, 99+% grade zirconyl chloride had Ag, Al, As, Cd, Cu, Hf, Mg, Na, Sc and V at 20, 1.8, 0.6, 0.6, 0.4, 8.4, 0.4, 2.4, 80 and 3 ppm, respectively. (See above.)

## METAL-ORGANIC COMPOUNDS

This section contains metal-organic compounds, ammonium and metal derivatives of organic alcohols, amines and carboxylic acids (salts), as well as ionophores that form complexes with metal ions. Note that there is a large number of metal-organic catalysts, and reagents for preparing some of these catalysts, in Chapter 6, and can be considered as an extension of this section. (For Introduction see p 555.)

**Acetylenedicarboxylic acid monopotassium salt** [928-04-1] **M 152.2**. It is very soluble in H<sub>2</sub>O, but can be crystallised from a small volume of H<sub>2</sub>O in small crystals. These are washed with EtOH and dried over H<sub>2</sub>SO<sub>4</sub> at 125°. [Bandrowski *Chem Ber* **10** 841 1877, Lossen *Justus Liebigs Ann Chem* **272** 133 1893, *Beilstein* **2** H 801, **2** I 317, **2** II 670, **2** III 1991, **2** IV 2290.]

**Acetylferrocene (ferrocenyl methylketone)** [1271-55-2] **M 228.1, m 86°, 86-87°**. Orange-red crystals are obtained when it is recrystallised from isooctane or \*C<sub>6</sub>H<sub>6</sub>, and then sublimed at 100°/1mm. The *oxime* has **m** 167-170° (from Et<sub>2</sub>O or aqueous EtOH). The *semicarbazone* has **m** 198-201° (from EtOH). [Richmond & Freiser *J Am Chem Soc* **77** 2022 1955, Weinmayer *J Am Chem Soc* **77** 3009 1955, Broadhead et al. *J Chem Soc* **650** 1958, *Beilstein* **16** IV 1798.]

**Allylpalladium(II) chloride dimer** [12012-95-2] **M 365.9, m ~160°**. It crystallises from benzene and is soluble in MeOH, Et<sub>2</sub>O and CHCl<sub>3</sub>. [Hüttel et al. *Chem Ber* **94** 766 1961, Dent et al. *J Chem Soc* 1585 1964, Armstrong *J Org Chem* **31** 618 1966.]

**Allyl tri-*n*-butylstannane (allyl tributyl tin)** [24850-33-7] **M 331.1, b 88-92°/0.2mm, 115°/17mm, d<sub>4</sub><sup>20</sup> 1.068, n<sub>D</sub><sup>20</sup> 1.487**. A possible impurity is tributylchlorostannane — test for Cl as Cl ion after hydrolysing. Dissolve it in \*C<sub>6</sub>H<sub>6</sub> (or toluene), shake this with dilute aqueous NaOH, dry (CaCl<sub>2</sub>), filter, evaporate and distil the residue in a vacuum [Jones et al. *J Chem Soc* 1446 1947, Bristow *Aldrichimica Acta* **17** 75 1984, Yamamoto *Aldrichimica Acta* **20** 45 1987]. [*Beilstein* **4** IV 4317.]

**Aluminum acetylacetonate (tris[2,4-pentandionate]aluminium)** [13963-57-0] **M 324.3, m 192-194°, 195°**. Recrystallise it several times from \*benzene or aqueous MeOH, λ<sub>max</sub> 216 and 286nm. [Charles & Pawlikowski *J Phys Chem* **62** 440 1958.] It can be purified by sublimation and has the following solubilities in g percent: \*C<sub>6</sub>H<sub>6</sub> 35.9 (20°), 47.6 (40°), toluene 15.9 (20°), 22.0 (40°) and acetylacetone 6.6 (20°), 10.4 (40°). [Fernelius & Bryant *Inorg Synth* **V** 105 1957, *Beilstein* **1** IV 3668.]

**Aluminum ethoxide** [555-75-9] **M 162.2, m 154-159°, 146-151°, b 187-190°/7mm, 210-214°/13mm**. Crystallise it from CS<sub>2</sub> [**m** 139°, CS<sub>2</sub> complex] and distil it in a vacuum. The molecular weight corresponds to [Al(OEt)<sub>3</sub>]<sub>4</sub> [Robinson & Peak *J Phys Chem* **39** 1127 1935, Vilani & Nord *J Am Chem Soc* **69** 2605 1947]. [*Beilstein* **1** H 313, **1** I 158, **1** II 3008, **1** III 1284, **1** IV 1289.]

**Aluminium isopropoxide** [555-31-7] **M 204.3, m 119°, b 94°/0.5mm, 135°/10mm**. Redistil it under vacuum. *Hygroscopic*. [Robinson & Peak *J Phys Chem* **39** 1127 1935, *Beilstein* **1** IV 1468.]

**Aluminum triethyl (triethyl aluminum)** [97-93-8] **M 114.2, b 69°/1.5mm, 76°/2.5mm, 129-131°/55mm, d<sub>4</sub><sup>20</sup> 0.695, n<sub>D</sub><sup>20</sup> 1.394**. Purify Al(Et)<sub>3</sub> by fractionation in an inert atmosphere under a vacuum in a 50cm column containing a heated nichrome spiral, taking the fraction **b** 112-114°/27mm. It is very sensitive to H<sub>2</sub>O and should be stored under N<sub>2</sub>. It should not contain chloride ions which can be shown by hydrolysis and testing with AgNO<sub>3</sub>. [Baker & Sisler *J Am Chem Soc* **75** 4828 5193 1953, NMR: Brownstein et al. *J Am Chem Soc* **81** 3826 1959, *Beilstein* **4** IV 4398.]

**Aluminium tri-*tert*-butoxide** [556-91-2] **M 246.3, m 208-210°(dec), >300°**. Crystallise Al(O-*t*Bu)<sub>3</sub> from \*C<sub>6</sub>H<sub>6</sub> and sublime it at 180°. [McElvain & Davie *J Am Chem Soc* **73** 1400 1951, *Beilstein* **1** IV 1612.]

**Aluminium trimethanide (trimethyl aluminium)** [75-24-1] **M 72.1, m 15.2°, b 111.5°/488.2mm, 124.5°/atm,  $d_4^{20}$  0.725.** Distil  $\text{Al}(\text{Me})_3$  through a 10-20 theoretical plates column under 1 atmosphere pressure of  $\text{N}_2$  (better with very slow take-off). It attacks grease (use glass joints). It has been distilled over Al in absence of grease, into small glass vials and sealed under  $\text{N}_2$ . The purity is measured by its freezing point. It reacts with  $\text{H}_2\text{O}$ , is non-conducting in  $\text{C}_6\text{H}_6$  and is **HIGHLY FLAMMABLE**. [Bamford et al. *J Chem Soc* 468 1946, Pitzer & Gutowsky *J Am Chem Soc* 68 2204 1946, *Beilstein* 4 IV 4397.]

**4-Aminophenylmercuric acetate** [6283-24-5] **M 371.8, m 168°, 175°(dec), 180°(dec).** Recrystallise it from hot dilute AcOH and dry it in air. Highly **TOXIC**. [Mahapatra et al. *J Indian Chem Soc* 32 613 1955, Albert & Schneider *Justus Liebigs Ann Chem* 465 269 1928, *Beilstein* 16 III 1411, 16 IV 1754.]

**Ammonium acetate** [631-61-8] **M 77.1, m 112-114°,  $d_4^{20}$  1.04.** Crystallise it twice from anhydrous acetic acid, and dry under vacuum for 24 hours at 100° [Proll & Sutcliffe *Trans Faraday Soc* 57 1078 1961].

**Ammonium benzoate** [1863-63-4] **M 139.2, m 198°, 200°(dec),  $d_4^{20}$  1.26.** Crystallise it from EtOH. [*Beilstein* 9 IV 273.]

**Ammonium dodecylsulfate (ammonium laurylsulfate)** [2235-54-3] **M 283.4.** Recrystallise it first from 90% EtOH and then twice from absolute EtOH, and finally dry it in a vacuum. [*Beilstein* 1 III 1786.]

**Ammonium ferric oxalate trihydrate** [13268-42-3] **M 428.1, m ~160°(dec),  $d_4^{20}$  1.77.** Crystallise it from hot water (0.5ml/g). [*Beilstein* 3 III 1103.]

**Ammonium formate** [540-69-2] **M 63.1, m 116°, 117.3°,  $d_4^{45}$  1.280.** Heat the solid in  $\text{NH}_3$  vapour and dry it in a vacuum till the  $\text{NH}_3$  odour is faint (note that it can evaporate completely in a vacuum). Recrystallise it from absolute EtOH and then keep it in a desiccator over 99%  $\text{H}_2\text{SO}_4$  *in vacuo*. It is very *hygroscopic*. It exists in two forms, stable needles and less stable plates. It also forms acid salts, i.e.  $\text{HCO}_2\text{NH}_4 \cdot 3\text{HCO}_2\text{H}$  and  $\text{HCO}_2\text{NH}_4 \cdot \text{HCO}_2\text{H}$ . [Kensall & Adler *J Am Chem Soc* 43 1473 1921, *Beilstein* 2 IV 18.]

**Ammonium ionophore I (Nonactin)** [6833-86-7] **M 736.9, m 147-148°,  $[\alpha]_D^{20}$  0° (c 1.2,  $\text{CHCl}_3$ ).** Crystallise it from MeOH (colourless needles), and it is dried at 20° in high a vacuum. It is a selectophore with high sensitivity for  $\text{NH}_4^+$  ions. [Corbaz et al. *Helv Chim Acta* 38 1445 1955, Domingues et al. *Helv Chim Acta* 45 129 1962, Nawata & Ando *Helv Chim Acta* 55 1371 1972, *Beilstein* 19/12 V 751.]

**Ammonium oxalate dihydrate** [6009-70-7] **M 142.1,  $d_4^{20}$  1.50.** Crystallise it from water (10ml/g) at 50°. [*Beilstein* 2 IV 1846.]

**Ammonium picrate** [131-74-8] **M 246.1, m EXPLODES above 200°.** Crystallise it from EtOH and acetone. [Mitchell & Bryant *J Am Chem Soc* 65 128 1943, *Beilstein* 6 II 262, 16 III 879, 16 IV 1392.]

***n*-Amylmercuric chloride** [544-15-0] **M 307.2, m 110°.** Crystallise it from EtOH. The *bromide* has **m 122°**. [Larock & Brown *J Am Chem Soc* 92 2467 1970, Marvel et al. *J Am Chem Soc* 47 3009 1925, *Beilstein* 14 H 706, 725.]

**Aurothioglucose (gold thioglucose)** [12192-57-3] **M 392.2.** Purify it by dissolving it in  $\text{H}_2\text{O}$  (0.05g in 1ml) and precipitating it by adding EtOH. It yields yellow crystals with a slight mercaptan odour. It decomposes slowly in  $\text{H}_2\text{O}$ , and is soluble in propylene glycol but insoluble in EtOH and other common organic solvents. [Cateron & Taylor *FEBS Lett* 98 351 1979, Cooney et al. *Biochem J* 259 651 1989.]

**Barium acetate** [543-80-6] **M 255.4.** Crystallise the salt twice from anhydrous acetic acid and dry it under vacuum for 24 hours at 100°. [*Beilstein* 2 I 49, 2 II 117, 2 III 192, 2 IV 114.]

**Barium ionophore I** [*N,N,N',N'*-tetracyclohexyloxy-bis-(*o*-phenyleneoxy)diacetamide] [96476-01-6] **M**

**644.9, m 156-158°.** Purify the ionophore by chromatography on a Kieselgel column, elute it with  $\text{CH}_2\text{Cl}_2/\text{EtOAc}$  (5:1), and recrystallise the residue from the evaporated effluent from  $\text{EtOH}/\text{Me}_2\text{CO}$  to give colourless crystals. It is an electrically neutral ionophore with high selectivity for  $\text{Ba}^{2+}$  ions and with high lipophilicity. [Kleiner et al. *Chem Ber* **118** 1071 1985, Läubli et al. *Anal Chem* **57** 2756 1985.]

**Barium propionate ( $\text{H}_2\text{O}$ )** [5908-77-0] **M 301.5,  $d_4^{27}$  1.44,  $\text{pK}^{24}$  4.88 (for propionic acid).** Crystallise it from warm water (50ml/g) by adding  $\text{EtOH}$  and cooling. [*Beilstein* **2** III 517, **2** IV 702.]

**Benzenechromium(0) tricarbonyl** [12082-08-5] **M 214.1, m 162-163°, 163-166°.** Purify the complex by sublimation *in vacuo*. A possible impurity is 2-picoline which can be removed by washing with pentane and drying. It is then purified further by sublimation at  $80\text{-}85^\circ/10^{-3}\text{mm}$ , or by recrystallisation from  $\text{Et}_2\text{O}$  to give yellow crystals.  $^1\text{H}$  NMR in  $\text{CDCl}_3$  should give a single peak at  $\tau$  4.68. [Rausch *J Org Chem* **39** 1787 1974, Pauson in Houben-Weyl *Meth Org Chem* V, E 18 Pt I p226 Theme Verlag, Stuttgart 1986, *Beilstein* **5** IV 625.]

**Beryllium acetate [ $\text{Be}(\text{OAc})_2$ ]** [543-81-7] **M 127.1, m 65-100° (slow heating), 155-180° (rapid heating).** It is obtained by dissolving the *basic acetate* (4g, see below) in boiling glacial  $\text{AcOH}$  containing acetyl chloride (4.5g) and refluxing for ~15 minutes. The  $\text{Be}(\text{OAc})_2$  which separates during this time, is filtered off, washed with  $\text{AcOH}$ , cold  $\text{CHCl}_3$  and dried *in vacuo* to give 90-94% yield of the salt. It is stable for several weeks in a tightly stoppered container at  $\sim 25^\circ$ . It slowly loses  $\text{Ac}_2\text{O}$ , and more rapidly on heating, to give the basic acetate which sublimes out. Strong heating partially decomposes it to give  $\text{Ac}_2\text{O}$  and  $\text{BeO}$ . It is not readily attacked by cold  $\text{H}_2\text{O}$  but forms a hydrate on warming, and is insoluble in most solvents. [Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 901 1963, Besson & Hardt *Z Anorg Allg Chem* **277** 188 1954.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium acetate (basic) [ $\text{Be}_4\text{O}(\text{OAc})_6$ ]** [1332-52-1] **M 406.3, m 283°, 285-286°, b 330-331°/atm.** On evaporating  $\text{Be}(\text{OH})_2$  with  $\text{AcOH}$  this covalent basic acetate is produced. *Alternatively*,  $\text{BeCO}_3$  (40g, see [744998-97-8]) and  $\text{AcOH}$  (80ml) are heated with stirring until evolution of  $\text{CO}_2$  ceases. When the reaction is complete, white semi-translucent crystals separate. After cooling to  $\sim 25^\circ$ , the basic acetate crystals are filtered off and dried in air. These are treated with warm  $\text{CHCl}_3$  (60-80ml), any insoluble material is filtered off, and the colourless octahedral crystals of the basic acetate which separate on cooling (with partial evaporation) are collected and freed from any  $\text{CHCl}_3$  *in vacuo* to give pure basic salt (28g, ~47%, **m 284°**). It is readily volatile, it can be distilled, and it sublimes in a vacuum leaving a small residue of 0.3-0.5% of  $\text{BeO}$ . It is very soluble in  $\text{CHCl}_3$ , soluble in non-polar organic solvents such as boiling  $^*\text{C}_6\text{H}_6$ , toluene, xylene, tetralin, and in  $\text{AcOH}$ ; less so in  $\text{CCl}_4$ ,  $\text{Ac}_2\text{O}$  and  $\text{AcCl}$ , but is sparingly soluble ( $\sim 0.3\%$ ) in  $\text{Et}_2\text{O}$ . It dissolves in anhydrous boiling  $\text{MeOH}$  but soon liberates  $\text{Ac}_2\text{O}$  to form highly aggregated basic acetates. It is quite stable in cold  $\text{H}_2\text{O}$ , but is rapidly hydrolysed in hot  $\text{H}_2\text{O}$ . [Moeller et al. *Inorg Synth* **III** 9 1950, Ehrlich in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 901 1963, and for the basic formate or propionate salts see Besson & Hardt *Z Anorg Allg Chem* **277** 188 1954, Hendus & Hardt *Z Anorg Allg Chem* **277** 127 1954, Hendus & Hardt *Z Anorg Allg Chem* **286** 265 1956, and Hardt *Z Anorg Allg Chem* **286** 254 1956, *Beilstein* **2** H 111, **2** I 48, **2** II 116, **2** III 190, **2** IV 112.] The X-ray crystal structure shows that the four beryllium atoms are arranged at the corners of a regular tetrahedron with oxygen at the centre, and the six edges are occupied by acetate groups. Thus the four Be atoms at the apexes of the tetrahedron are coordinated to the carboxyl oxygen atoms of the six acetate groups, i.e. each Be atom is coordinated to three oxygen atoms. [Bragg & Morgan *Proc Roy Soc* **104** 437 1923, for partial structure see also F.A. Cotton and G. Wilkinson *Advanced Inorganic Chemistry (A Comprehensive Text)* p. 175, Interscience Publ. 1962, Library of Congress Catalog Number 62-14818.] Beryllium compounds are potentially **carcinogenic**.

**Beryllium acetylacetonate [ $\text{Be}(\text{acac})_2$ ]** [10210-64-7] **M 207.2, m 108°, 108.5-109°, b 270°,  $d^{20}$  1.168.** This complex has been prepared by adding 3M aqueous  $\text{NaOH}$  (100ml) to a solution of  $\text{BeSO}_4 \cdot 4\text{H}_2\text{O}$  (10g, see [7787-50-0]) in distilled  $\text{H}_2\text{O}$  (100ml) to give a strong basic solution to which is added freshly distilled acetylacetone (30ml) dropwise with stirring, and a white precipitate separates. This is collected, washed with  $\text{H}_2\text{O}$  and dried *in vacuo* to give the complex (5.1g 43%) which is recrystallised by dissolving in a small volume of  $^*\text{C}_6\text{H}_6$  and adding gradually petroleum ether until crystallisation is complete. The crystals are collected, washed with

petroleum ether and dried *in vacuo*. It can be recrystallised further from a small volume of \*C<sub>6</sub>H<sub>6</sub> or a large volume of petroleum ether. It has also been prepared in 70% yield from basic beryllium carbonate (3g) in H<sub>2</sub>O (45ml) by conversion to the chloride with 6N aqueous HCl (*ca* 20ml) until slightly acidic. To this is added dropwise a clear solution of acetylacetone (10g suspended in 45ml of H<sub>2</sub>O to which is added 6N ammonia until completely clear) and stirred until precipitation of the complex is complete (pH should be near neutral). The solid is collected washed with H<sub>2</sub>O, dried, and crystallised as before. It is also sublimed through a plug of glass wool at a vacuum of 0.1mm by heating in a bath at 80 to 100°. It is very soluble in most organic solvents such as EtOH, Et<sub>2</sub>O, Me<sub>2</sub>CO, \*C<sub>6</sub>H<sub>6</sub> and in CS<sub>2</sub> but considerably less in petroleum ethers. However, it is practically insoluble in H<sub>2</sub>O which hydrolyses it on boiling; and it is also hydrolysed by acids and alkalis. [Jones *J Am Chem Soc* **81**, 3188 1959, Arch & Young *Org Synth II* 17 1946.] Beryllium compounds are potentially carcinogenic.

**Bicyclo[2.2.1]hepta-2,5-diene rhodium (I) chloride dimer (norbornadiene rhodium chloride complex dimer)** [12257-42-0] **M 462, m 240°(dec)**. It recrystallises from hot CHCl<sub>3</sub>/petroleum ether as fine crystals soluble in CHCl<sub>3</sub> and \*C<sub>6</sub>H<sub>6</sub> but is almost insoluble in Et<sub>2</sub>O or petroleum ether. [Abel et al. *J Chem Soc* 3178 1959.]

**2,2'-Bipyridinium chlorochromate** [76899-34-8] **M 292.6**. Wash it with cold conc HCl, then H<sub>2</sub>O (sintered glass funnel) and dry it in a vacuum (CaCl<sub>2</sub>) to give a free-flowing yellow-brown powder. Store it in the dark. [Guziec & Luzzio *Synthesis* 691 1980, Chakraborty & Chandrasekaran *Synth Commun* **10** 951 1980.] **SUSPECTED CARCINOGEN.**

**2,2'-Biquinolin-4,4'-dicarboxylic acid dipotassium salt** [63451-34-3] **M 420.51**. Recrystallise it from H<sub>2</sub>O. The *Cu salt* has UV with  $\lambda_{\max}$  at 562nm. [Mopper & Gindler *Anal Biochem* **56** 440 1973, Beilstein **25** IV 1148.]

**Bis(1,2,3,4,5- $\eta$ -cyclooctadienyl)ruthenium [Ru(COD)<sub>2</sub>]** [63395-36-8] **M 315.4, m 89-90°, 113-114°**. The bis-complex is prepared by adding a solution of RuCl<sub>3</sub> · 3H<sub>2</sub>O (624mg, 1.33mol) in EtOH (20ml) under argon (or N<sub>2</sub>) slowly over 30 minutes to a stirred suspension of Zn powder (7g), EtOH (5ml) and redistilled 1,2-5,6-cyclooctadiene (11ml, 80mol), COD see [111-78-4, 1552-12-1], then stirring at room temperature for 2 hours, and filtering the mixture under N<sub>2</sub>. The filtrate is concentrated on a vacuum line to remove volatiles, and the residual brown oil is placed on top of an Al<sub>2</sub>O<sub>3</sub> column in hexane after flushing with N<sub>2</sub>. The yellow band is eluted with hexane, concentrated *in vacuo*, the eluate is evaporated, pentane is added to dissolve the residue and the solution is cooled at -78° for several hours, to provide very pale yellow (almost colourless) crystals of the analytically pure complex (283mg, 35%). The <sup>1</sup>H NMR (60MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has  $\tau$  at 4.42 (t, 2H), 6.15 (dd, 4H), 6.45 (m, 4H) and 7.49-8.95 (m, 12H); and the <sup>13</sup>C NMR (22.6MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has  $\delta$  at 22.5, 27.3, 36.7, 35.3, 38.9, 62.4, 76.7, 87.6. [Pertici et al. *J Chem Soc, Perkin Trans* 1961 1980 (gave m 89-90°), Itoh et al. *J Organomet Chem* **272** 179 1984 (gave m 113-114°)].

**Bis( $\eta^5$ -2,4-cyclopentadien-1-yl)samarium II [SmCp<sub>2</sub>, samarocene]** [80695-16-5] **M 280.5, decomposes at high temperatures**. Samarium dicyclopentadienyl can be prepared in Schlenk equipment under argon by adding SmI<sub>2</sub> (0.1M in THF, 60ml, 6mmol) to a 0.4M solution of sodium pentadienide in THF (30ml, 12mmol) when the complex separates immediately and is decanted within 1 hour. The red powder of SmCp<sub>2</sub> is collected, washed twice with THF to remove NaI, dried *in vacuo* and stored under argon. The red powder and the purple solid THF complex (**M 352.6**) are *pyrophoric* in air. SmCp<sub>2</sub> can also be stored for a few days in a Schlenk tube under THF without decomposition. Freshly prepared suspensions of SmCp<sub>2</sub> have been used in pseudo-Barbier couplings between carbonyl compounds and aliphatic or allylic halides and are more efficient than SmI<sub>2</sub> [Namy et al. *J Organometal Chem* **328** 81 1986].

**SmCp<sub>2</sub>-THF** has been prepared in one arm of a U-tube by stirring in an anhydrous oxygen-free helium atmosphere) tris(cyclopentadienyl)samarium III (SmCp<sub>3</sub>, 0.499g), K metal (0.527g, naphthalene (0.128g, 25% molar deficiency) and THF (25ml, freshly distilled under reduced pressure from Na-benzophenone) for 48 hours. Purple SmCp<sub>2</sub>-furanate was separated by decantation, washed by successive back-distillation with THF and the

solvent is removed at 24°/10<sup>-3</sup>mm during 18 hours. It is *pyrophoric* and should be handled in an anhydrous oxygen-free atmosphere. Its IR has  $\nu_{\max}$  (Nujol) at 3080 (C-H stretch), 1475, 1347, 1308, 1263, 1163, 1070, 1008 (C-H bend parallel), 775, 740 (C-H bend perpendicular) and 350 (antisym metal-ring vibration) cm<sup>-1</sup>, and other bands at 2980, 2880, 1375, 725 and 565 cm<sup>-1</sup> due to coordinated THF with intensities similar to the weaker metallocene bands. On exposure to traces of air this complex immediately changes in colour from deep purple to yellow-gray with drastic reduction in the paramagnetism. [Watt & Gillow *J Am Chem Soc* **91** 775 1969.]

**Bis(2,9-dimethyl-1,10-phenanthroline) copper(I) perchlorate (Cuproine)** [54816-44-5] **M 579.6, pK<sup>25</sup> -2.4 to -3.1 (for HClO<sub>4</sub>) and 6.15 (for dimethylphenanthroline).** Crystallise it from acetone. It has UV with  $\lambda_{\max}$  in isopentanol or hexan-1-ol at 454nm. [Smith & McCurdy *Anal Chem* **24** 371 1952, Gahler *Anal Chem* **26** 577 1954, *Beilstein* **23** III/IV 1737.]

**1,1'-Bis(diphenylphosphino)ferrocene** [12150-46-8] **M 554.4, m 181-183°, 184-194°.** Wash the ferrocene with distilled H<sub>2</sub>O and dry it in a vacuum. Dissolve it in *ca* 5 parts of hot dioxane and cool to give orange crystals **m** 181-183°. Recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/heptane (1:2) gives a product with **m** 183-184°. [Bishop et al. *J Organomet Chem* **27** 241 1971.]

**Bis(ethyl)titanium(IV) chloride** [2247-00-9] **M 177.0.** Recrystallise it from boiling toluene. [See Marek (ed) *Titanium and Zirconium in Organic Synthesis* Wiley-VCH 2002 ISBN 3-527-304-28-2, Wailes et al. *Organometallic Chemistry of Titanium, Zirconium and Hafnium* Academic Press 1974 ISBN 0-127303502.]

**Bis(ethyl)zirconium(IV) chloride** [92212-70-9] **M 220.3.** Recrystallise it from boiling toluene. [See Marek (ed) *Titanium and Zirconium in Organic Synthesis* Wiley-VCH 2002 ISBN 3-527-304-28-2, Wailes et al. *Organometallic Chemistry of Titanium, Zirconium and Hafnium* Academic Press 1874 ISBN 0-127303502.]

**N,N'-Bis(salicylidene)ethylenediamine cobalt (II) [Co(SALEN)<sub>2</sub>, salcomine]** [14167-18-1] **M 325.2.** The powder should have an oxygen capacity of 4.7-4.8% as measured by the increase in weight under O<sub>2</sub> at 100 pounds pressure at *ca* 20°. The O<sub>2</sub> is expelled on heating the material to 65°. It crystallises from pyridine, CHCl<sub>3</sub> or \*C<sub>6</sub>H<sub>6</sub>, and the solvent may be removed by heating at 120° in a vacuum. However, this heating may mean reduced O<sub>2</sub> capacity. In the dry state it absorbs O<sub>2</sub>, turning from a maroon colour to black. [Diehl & Hack *Inorg Synth* **III** 196 1950.]

**Bis(tetrabutylammonium) dichromate** [56660-19-6] **M 700.9, m 139-142°.** Wash the dichromate with water and dry it in a vacuum. Recrystallise it from hexane (**m** 79-80°). [Santaniello & Ferraboschi *Synth Commun* **10** 75 1980, *Beilstein* **4** IV 556.] (Possible **CARCINOGEN**.)

**n-Butylmercuric chloride** [543-63-5] **M 293.1, m 130°.** Crystallise it from 95% EtOH. [Larock & Brown *J Am Chem Soc* **92** 2467 1970, Marvel et al. *J Am Chem Soc* **47** 3009 1925.]

**n-Butylstannoic acid [PhSn(OH)<sub>3</sub>, trihydroxy-n-butylstannane]** [22719-01-3] **M 208.8.** Purify it by adding excess KOH in CHCl<sub>3</sub> to remove *n*-BuSn(OH)Cl<sub>2</sub> and *n*-BuSn(OH)<sub>2</sub>Cl, and isolate it by acidification [Holmes et al. *J Am Chem Soc* **109** 1408 1987].

**Cadion [1-(4-nitrophenyl)-3-(4-phenylazophenyl)-triazene]** [5392-67-6] **M 346, m 189°(dec).** Commercial cadion is purified by recrystallisation from 95% EtOH and is dried *in vacuo*. It is stable in 0.2 N KOH (in 20% aqueous EtOH) at 25°. It is a sensitive reagent for Cd, and the Cd complex has UV with  $\lambda_{\max}$  (EtOH) at 475nm. [Chavanne & Geronimi *Anal Chim Acta* **19** 377 1958, *Beilstein* **16** III 664.]

**Cadmium acetate dihydrate** [5743-04-4] **M 230.5, m 255°(anhydrous), d<sub>4</sub><sup>20</sup> 2.01 (hydrate), 2.34 (anhydr), pK<sub>1</sub><sup>25</sup> 9.7, pK<sub>2</sub><sup>25</sup> ~11.0 (for Cd<sup>2+</sup>).** Recrystallise it twice from anhydrous acetic acid and dry it under vacuum for 24 hours at 100°. [*Beilstein* **2** IV 114.]

**Cadmium ionophore I [N,N,N',N'-tetramethyl-3,6-dioxooctanedi-(thioamide)]** [73487-00-0] **M 432.7, m**

**35-36°.** Wash it well with petroleum ether, then several times with 2N HCl (if it has a slight odour of pyridine), then H<sub>2</sub>O and dry it in a vacuum over H<sub>2</sub>SO<sub>4</sub>. It is a polar selectrophore for Cd. [Schneider *Helv Chim Acta* **63** 217 1980, Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Cadmium lactate** [16039-55-7] **M 290.6.** Recrystallise the lactate from water (10ml/g) by partial evaporation in a desiccator. [Beilstein **3** H 277, **3** III 465, **2** IV 637.]

**Cadmium salicylate** [19010-79-8] **M 248.5, 242°(dec).** Recrystallise it from distilled H<sub>2</sub>O by evaporation in a desiccator. It is an antiseptic. [Prasad et al. *J Indian Chem Soc* **35** 267 1958, *Beilstein* **10** H 60, **10** I 25, **10** II 33, **10** III 94, **10** IV 128.]

**Calcein sodium salt** [2',7'-bis-{N,N-di(carboxymethyl)aminomethyl}fluorescein Na salt, **Fluorexon, Fluorescein Complexon**] [108750-13-6 diNa salt, 1461-15-0 free acid] **M 666.5, pK<sub>Est(1)</sub>~1.9, pK<sub>Est(2)</sub>~2.5, pK<sub>Est(3)</sub>~8.0, pK<sub>Est(4)</sub>~10.5 (all for N-CH<sub>2</sub>COOH), and pK<sub>Est(5)</sub>~3.5 (for benzoic COOH).** Dissolve it in distilled H<sub>2</sub>O and acidify with dilute HCl to pH 3.5. Filter off the solid acid and wash it well with H<sub>2</sub>O. Redissolve *ca* 10g in 300ml H<sub>2</sub>O containing 12g of NaOAc. Precipitate it again by adding HCl, filtering and washing with H<sub>2</sub>O. Add the solid to 200ml of EtOH stir for 1 hour and filter. Repeat the EtOH wash and dry the bright yellow solid in a vacuum. This acid decomposes on heating at *ca* 180°. See below for the preparation of the Na salt. [Diehl & Ellingboe *Anal Chem* **28** 882 1956].

*Alternatively*, dissolve it in H<sub>2</sub>O and acidify with 3N HCl to pH 3.5. Collect the solid and wash it with H<sub>2</sub>O. The air-dried precipitate is extracted with 70% aqueous EtOH, filtered hot and cooled slowly. Fine yellow needles of the acid crystallise out; they are filtered off and dissolved in the minimum quantity of 0.01N NaOH and re-precipitated by adding N HCl to pH 3.5. It is then recrystallised from 70% aqueous EtOH (3x). The final product (acid) is dried at 80° in a vacuum for 24 hours, **m >300°(dec)**. It contains one molecule of water per molecule of acid (C<sub>30</sub>H<sub>36</sub>N<sub>4</sub>O<sub>13</sub>·H<sub>2</sub>O). The product is pure as revealed by electrophoresis at pH 5.6 and 8.6, and by TLC in *i*-BuOH/*i*-PrOH/AcOH/H<sub>2</sub>O (60:60:5:5 by volume) or *i*-PrOH or pH 8.0 borate buffer. [Wallach et al. *Anal Chem* **31** 456 1959.]

The Na salt is prepared by dissolving the pure acid in H<sub>2</sub>O containing 2 mols of NaOH per mol of acid reagent and lyophilising. It complexes with Ca and Mg ions. [Beilstein **19** III/IV 4338.]

**Calcium acetate monohydrate** [5743-26-0 (H<sub>2</sub>O), 62-54-4 (xH<sub>2</sub>O)] **M 176.2 (H<sub>2</sub>O), m 150° (loses H<sub>2</sub>O), pK<sup>25</sup> 12.7 (for Ca<sub>2</sub><sup>+</sup>).** Recrystallise the acetate from water (3ml/g) by partial evaporation in a desiccator. [Beilstein **2** IV 113.]

**Calcium benzoate trihydrate** [2090-05-3] **M 336.4.** Recrystallise the benzoate from water (10ml/g) between 90° and 0°. [Beilstein **9** I 60, **9** H 85, **9** III 377, **9** IV 280.]

**Calcium butyrate** [5743-36-2] **M 248.2, d<sub>4</sub><sup>30</sup> 1.271.** Recrystallise the butyrate from water (5ml/g) by partial evaporation in a desiccator and dry it in a vacuum to constant weight. [Pathak & Bhide *J Indian Chem Soc* **30** 47, 48 1953.] Its dissociation constant at 25° is 0.29 [Colman-Porter & Monk *J Chem Soc* 4363 1952, *Beilstein* **2** IV 785].

**Calcium carbamate** [543-88-4] **M 160.1.** Recrystallise calcium carbamate from aqueous ethanol. [Beilstein **4** H 75, **4** I 336, **4** II 557, **4** III 149, **4** 234.]

**Calcium formate** [544-17-2] **M 130.1, m dec on heating, d<sub>4</sub><sup>20</sup> 2.01.** Recrystallise calcium formate from water (5ml/g) by partial evaporation in a desiccator. [Beilstein **2** IV 16.]

**Calcium D-gluconate monohydrate** [299-28-5, 18016-24-5] **M 448.4, m dec on heating, [α]<sub>D</sub><sup>20</sup> +11.0°, [α]<sub>D</sub><sup>20</sup> +9.0° (c 1.2, H<sub>2</sub>O).** Calcium gluconate is soluble in H<sub>2</sub>O (3.5g in 100g at 25°). Dissolve it in H<sub>2</sub>O, filter and precipitate it by adding MeOH. Filter off the solid and dry it in a vacuum at 85°. *Alternatively*, dissolve it in H<sub>2</sub>O, filter (from insoluble inorganic Ca) and evaporate it to dryness under vacuum at 85°. [March et al. *J Am Pharm Assoc* **41** 366 1952, *Beilstein* **3** IV 1255.]

**Calcium D-heptagluconate dihydrate** [17140-60-2] **M 526.4**,  $[\alpha]_{546}^{20} +5.2^{\circ}$ ,  $[\alpha]_{\text{D}}^{20} +4.4^{\circ}$  (c 5, H<sub>2</sub>O). Purify it in the same way as for calcium D-gluconate. [*Beilstein 3* III 1112.]

**Calcium ionophore I (ETH 1001)** [58801-34-6] **M 685.0**. This is a neutral Ca selectophore. It can be purified by thick layer (2mm) chromatography (Kieselgel F<sub>245</sub>) and eluted with Me<sub>2</sub>CO/CHCl<sub>3</sub> (2:1). [Ammann et al. *Helv Chim Acta* **56** 1780 1973, Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Calcium ionophore II (ETH 129)** [74267-27-9] **M 460.7**, **m 153-154<sup>o</sup>**. Recrystallise it from Me<sub>2</sub>CO. It forms 1:2 and 1:3 metal/ligand complexes with Mg<sup>2+</sup> and Ca<sup>2+</sup> ions, respectively, and induces selectivity in membranes for Ca<sup>2+</sup> over Mg<sup>2+</sup> by a factor of *ca* 10<sup>4</sup>. [Pretsch et al. *Helv Chim Acta* **63** 191 1980, Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Calcium ionophore III [A23187 calcimycin]** [52665-69-7] **M 523.6**, **m 181-182<sup>o</sup>**,  $[\alpha]_{\text{D}}^{25} -56.0^{\circ}$  (c 1, CHCl<sub>3</sub>). It recrystallises from Me<sub>2</sub>CO as colourless needles. Protect it from light and moisture, store in a refrigerator. It is soluble in Me<sub>2</sub>SO or EtOH and can be stored for 3 months without loss of activity. The Mg and Ca salts are soluble in organic solvents and cross biological membranes. It has a pK<sub>a</sub> of 6.9 in 90% Me<sub>2</sub>SO. The Ca complex crystallises from 50% EtOH as colourless prisms. It is *highly TOXIC*. [Pressman *Ann Rev Biochem* **45** 501 1976, Chaney et al. *J Am Chem Soc* **96** 1932 1974, Chaney et al. *J Antibiotics* **29** 124 1976, Suzuki et al. *Anal Biochem* **61** 382 1989, Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Calcium isobutyrate** [533-90-4] **M 248.2**. Crystallise it from water (3ml/g) by partial evaporation in a desiccator. It forms a *pentahydrate* at low temperatures, but the crystals filtered from a saturated solution at 80<sup>o</sup> are the *monohydrate*; the transition temperature is 62.5<sup>o</sup>. [Lumsden *J Chem Soc* **81** 359 1902.] It has a dissociation constant of 0.31 [Colman-Porter & Monk *J Chem Soc* 4363 1952]. [*Beilstein 2* H 290, **2** II 290, **2** IV 845.]

**Calcium lactate (5H<sub>2</sub>O)** [814-80-2, 15743-47-5] **M 308.3**, **m anhydrous at 120<sup>o</sup>**,  $[\alpha]_{\text{D}}^{20} -4.2^{\circ}$  (c 5, H<sub>2</sub>O). Crystallise calcium lactate from warm water (10ml/g) by cooling to 0<sup>o</sup>. [*Beilstein 3* IV 636.]

**Calcium propionate** [4075-81-4] **M 186.2**, **m dec on heating**. Crystallise this antifungal salt from water (2ml/g) by partial evaporation in a desiccator. [*Beilstein 2* H 238, **2** II 218, **2** III 516.]

**Calcium salicylate (2H<sub>2</sub>O)** [824-35-1] **M 350.4**, **pK<sub>1</sub><sup>20</sup> 3.08**, **pK<sub>2</sub><sup>20</sup> 13.43 (for acid)**. Recrystallise calcium salicylate from water (3ml/g) between 90<sup>o</sup> and 0<sup>o</sup>. [*Beilstein 10* H 60, **10** II 33, **10** III 94, **10** IV 128.]

**Carbonate ionophore I [ETH 6010] (heptyl 4-trifluoroacetylbenzoate)** [129476-47-7] **M 316.3**, **b 170<sup>o</sup>/0.02mm**, **d<sub>4</sub><sup>20</sup> 0.909**. Purify the ionophore by flash chromatography (2g of reagent with 30g of Silica Gel 60) and elute with EtOAc/hexane (1:19). The fractions that absorb light at 260nm are pooled, evaporated and dried at room temperature (10.3 Torr). The oily residue is distilled in a bubble-tube apparatus (170<sup>o</sup>/0.02 Torr). Its IR (CHCl<sub>3</sub>) has peaks at 1720, 1280, 940cm<sup>-1</sup>, and its solubility in tetrahydrofuran is 50mg/0.5ml. It is a lipophilic neutral ionophore selective for carbonate as well as being an optical humidity sensor. [Behringer et al. *Anal Chim Acta* **233** 41 1990.]

**Cerous acetate** [537-00-8] **M 317.3**, **pK<sub>1</sub><sup>25</sup> 8.1 (9.29)**, **pK<sub>2</sub><sup>25</sup> 16.3**, **pK<sub>3</sub><sup>25</sup> 26.0 (for Ce<sup>3+</sup>)**. Recrystallise it twice from anhydrous acetic acid, then pumped dry under a vacuum at 100<sup>o</sup> for 8 hours. [*Beilstein 2* I 50, **2** II 119, **2** III 196, **2** IV 115.]

**Cesium oleate** [31642-12-3] **M 414.4**. Recrystallise cesium oleate from EtOAc, dry it in an oven at 40<sup>o</sup> and store it over P<sub>2</sub>O<sub>5</sub>. [Finkle et al. *J Am Chem Soc* **45** 2785 I1923, *Beilstein 2* II 437, **2** III 1405.]

**Cesium perfluoro-octanoate (Cesium pentadecafluorooctanoate)** [17125-60-9] **M 546.0**. Recrystallise it from a butanol/petroleum ether mixture, dry it in an oven at 40<sup>o</sup> and store it over P<sub>2</sub>O<sub>5</sub> under vacuum. [*Beilstein 2* IV 994.]

**Chloro-tri-isopropoxy titanium** [20717-86-6] **M 260.6, m 45-50°, b 61-65°/0.1mm**. When distilled under vacuum, the distillate sets slowly to a solid on standing. Stock reagents are made by dissolving the warm liquid in pentane, toluene, Et<sub>2</sub>O, THF, CH<sub>2</sub>Cl<sub>2</sub>, and can be stored in a pure state or in solution under dry N<sub>2</sub> for several months. The reagent is *hygroscopic* and is hydrolysed by H<sub>2</sub>O. [Reetz et al. *Chem Ber* **118** 1421 1985.]

**Chromocene [bis(cyclopentadienyl) chromium II]** [1271-24-5] **M 182.2, m 173°**. Chromocene forms *pyrophoric* red crystals on sublimation at 50°/0.1mm followed by resublimation at 75-90°/0.1mm. Although it is stable at least to 300°, it is readily oxidised in air, and effervesces slowly in H<sub>2</sub>O to give cyclopentadiene. All operations should be carried out in a dry box. It decomposes in CCl<sub>4</sub> or CS<sub>2</sub>, and for IR even grinding with nujol, KBr or KI causes some decomposition. [Wilkinson et al. *J Inorg Nucl Chem* **95** 109 1956, Wilkinson *J Am Chem Soc* **76** 209 1954, *Beilstein* **16** IV 1774.]

**Chromoionophore I [ETH 5294] [9-diethylamino-5-octadecanoyl-imino-5-H-benzo[a]-phenoxazine]** [125829-24-5] **M 583.9**. Purify it by flash chromatography (Silica Gel) and elute with EtOAc. The coloured fractions are pooled, evaporated and recrystallised from EtOAc. It is a lipophilic fluorescent chromoionophore and is a selectophore for K and Ca ions. [Morf et al. *Anal Chem* **62** 738 1990.]

**Cobalt (II) meso-5,10,15,20-tetraphenylporphine complex** [14172-90-8] **M 671.7**. It yields brown crystals from Et<sub>2</sub>O or CHCl<sub>3</sub>/MeOH (*cf iron chloride complex*). It crystallises on extraction (Soxhlet) with \*C<sub>6</sub>H<sub>6</sub>. It is soluble in most organic solvents except MeOH and petroleum ether. [UV, IR: Rothmund & Manott *J Am Chem Soc* **70** 1808 1948, Thomas & Martell *J Am Chem Soc* **81** 5111 1959.]

**Cobaltic acetylacetonate** [21679-46-9] **M 356.3**. Recrystallise it from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and dry it in a vacuum. [Charles & Pawlikowski *J Phys Chem* **62** 440 1938, *Beilstein* **1** H 783.]

**Cobaltous acetate tetrahydrate** [6147-53-1] **M 249.1, pK<sub>1</sub><sup>25</sup> 9.85 (for Co<sup>2+</sup>)**. Several recrystallisations from 50% aqueous acetic acid give the *tetrahydrate*. It is converted to the *anhydrous* salt by drying at 80°/1mm for 60 hours. [*Beilstein* **2** IV 120.]

**Cobaltous acetylacetonate** [14024-48-7, 123334-29-2 x H<sub>2</sub>O] **M 257.2, m 165-170°, 172°**. Recrystallise the complex from acetone or MeOH and dry it in a vacuum. [*Beilstein* **1** H 783.]

**12-Crown-4 (lithium ionophore V, 1,4,7,10-tetraoxacyclododecane)** [294-93-9] **M 176.2, m 17°**. The distilled crude ionophore has to be recrystallised from pentane at -20° to remove acyclic material. It is then dried over P<sub>2</sub>O<sub>5</sub>. It complexes Li. [Anet et al. *Acta Chem Scand* **27** 3395 1973, *Beilstein* **19/11** V 334.]

**Cupric acetate monohydrate** [6046-93-1 (H<sub>2</sub>O), 142-71-2 (anhydrous)] **M 199.7, m 115°, 240°(dec), d<sub>4</sub><sup>20</sup> 1.88, pK<sub>1</sub><sup>25</sup> 8.0, pK<sub>2</sub><sup>25</sup> 13.1 (for Cu<sup>2+</sup>)**. Recrystallise Cu(OAc)<sub>2</sub> twice from warm dilute acetic acid solutions (5ml/g) by cooling. [*Beilstein* **2** IV 111.]

**Cupric benzoate** [533-01-7] **M 305.8**. Recrystallise it from hot water. Its solubility in EtOH/\*C<sub>6</sub>H<sub>6</sub> (90%) at 25° is 0.1%. [Crawford & Stewart *J Chem Soc* 228, 289 1953, *Beilstein* **9** H 84, **9** I 60, **9** III 376, **9** IV 280.]

**Cupric lactate monohydrate** [814-81-3] **M 295.7**. The *monohydrate* crystallises from hot H<sub>2</sub>O (3ml/g) on cooling. [*Beilstein* **3** II 203, **3** III 465, **3** IV 636.]

**Cupric oleate** [1120-44-1] **M 626.5**. Crystallise cupric oleate from diethyl ether. [*Beilstein* **2** H 465, **2** I 196, **2** I 202, **2** II 436, **2** III 1404, **2** IV 1646.]

**Cupric phthalocyanine** [147-14-8] **M 576.1**. Precipitate it twice from conc H<sub>2</sub>SO<sub>4</sub> by slow dilution with water. It has also been purified by two or three sublimations at 580° in an argon flow at 300-400Pa. [*Beilstein* **26** III/IV 4256.]

**(1,5-Cyclooctadiene)(1,3,5-cyclooctatetraene)ruthenium [Ru(COD)(COT)]** [127382-91-6] **M 315.4, m 88-**

**94°**, **92-94°**. It can be prepared directly under argon and pressure equalisation, by adding freshly distilled cycloocta-1,5-diene (12.5ml, COD, see [111-78-4, 1552-12-1]) in MeOH (5ml) to Zn dust (6g) in a flask which is placed in an ultrasound bath (Branson, Bransonic Model B220) thermostated at 70°. A solution of RuCl<sub>3</sub>·3H<sub>2</sub>O (530mg) in MeOH (12ml) is slowly added during 20 minutes to the refluxing COD mixture under N<sub>2</sub> flow and ultrasound. After addition is complete the mixture is sonicated for a further 2 hours under reflux at 70°. The solids are filtered off and the filtrate is evaporated with a vacuum line. The residual brown oil is extracted with the minimum volumes of hexane (3 x, total ~70ml), placed on an Al<sub>2</sub>O<sub>3</sub> (Merck 1079, or Brockman Activity II-III) column under N<sub>2</sub> and eluted with hexane. The yellow band is collected, the volatiles are removed *in vacuo* to give orange crystals (590mg, 93%). Analytically pure **Ru(COD)(COT)** [C<sub>16</sub>H<sub>22</sub>Ru] is obtained as yellow crystals upon recrystallisation (70-90% recovery) from hexane at ~ -78°/-6 hours preferably under N<sub>2</sub> or argon. [Itoh et al. *J Organomet Chem* **272** 179 1984, Pertici et al. *J Chem Soc, Dalton Trans* 1961 1980.] The <sup>1</sup>H NMR (60MHz, \*C<sub>6</sub>D<sub>6</sub>, TMS) has τ at 4.78 (dd, 2H), 5.22 (m, 2H), 6.21 (m, 2H), 7.08 (m, 4H), 7.78 (m, 8H), 8.36 (m, 2H) and 9.10 (m, 2H); compare with the <sup>1</sup>H NMR (60MHz, \*C<sub>6</sub>D<sub>6</sub>, TMS) of **Ru(COD)<sub>2</sub>** (see below) which has τ at 4.42 (t, 2H), 6.15 (dd, 4H), 6.45 (m, 4H), and 7.49—8.95 (m, 12H).

**Alternatively**, thermal isomerisation of **Ru(COD)<sub>2</sub>** (75mg, see [63395-36-8]) in CDCl<sub>3</sub> (0.5ml) in an NMR tube filled with argon occurs when the sealed tube is heated at 70° for 12 minutes as observed by the spectrum. Evaporation of the solution and purification as above gave yellow-orange crystals of **Ru(COD)(COT)** in >85% yield with no deuterium incorporation as observed by NMR. Isomerisation in \*C<sub>6</sub>D<sub>6</sub> is very slow at 70° giving a mixture of three isomers after 5 hours [Itoh et al. *J Organomet Chem* **272** 179 1984, see also Petrici et al. *J Chem Soc, Dalton Trans* 1961 1980]. A synthesis *via* a Grignard reduction of [RuCl<sub>2</sub>(COD)]<sub>n</sub> has also been described [Pelzer et al. *Chem Mater* **16** 4937 2004]. For its catalytic activity see Chapter 7, Catalysts-Part 1. It is a key precursor for the preparation of stabilised Ru nanoparticles [Pan et al. *J Am Chem Soc* **123** 7584 2001, cf Chapter 8].

**Cyclopentadienyl iron(II) dicarbonyl dimer [di(cyclopentadienyl)tetracarbonyl-diiron, (C<sub>5</sub>H<sub>5</sub>Fe)<sub>2</sub>(CO)<sub>4</sub>] [38117-54-3] M 353.9, m 194° (dec) (sealed tube)**. This key precursor to a variety of cyclopentadienyl-carbonyl complexes is prepared from Fe(CO)<sub>5</sub> (10ml, [13463-40-6], **CARE**, poisonous, work in an efficient fume cupboard) by refluxing with an excess of dicyclopentadiene (10ml, cyclopentadiene dimer see 77-73-6) for *ca* 40 hours under a N<sub>2</sub> or argon atmosphere in the absence of light, removing the solvent, adding CHCl<sub>3</sub> (50ml), and removing any Fe<sub>2</sub>(CO)<sub>9</sub> and insoluble material by centrifugation. The decanted supernatant is cooled in a Dry-ice/Me<sub>2</sub>CO bath, the crystals are collected by centrifugation and recrystallised from CHCl<sub>3</sub> to give a ~30% yield (based on CO) of analytically pure dark reddish-purple crystals of (C<sub>5</sub>H<sub>5</sub>Fe)<sub>2</sub>(CO)<sub>4</sub>, **m 194° (dec)** (sealed tube). The solid is stable to air and light, but solutions deteriorate under these conditions. It forms ferrocene (in 75% yield based on C<sub>5</sub>H<sub>5</sub>) when heated at 210°. It is readily soluble in EtOH, CHCl<sub>3</sub>, pyridine, less soluble in CCl<sub>4</sub> and CS<sub>2</sub>, and sparingly soluble in light petroleum giving red solutions in each case. It is insoluble in H<sub>2</sub>O and is unaffected by it. The measured molecular weight by isothermal distillation is 368. The IR has strong bands at λ<sub>max</sub> at 4.95μ, 5.1μ and 5.6μ (in CCl<sub>4</sub>) and 12.1μ (CS<sub>2</sub>). In the presence of O<sub>2</sub> a solution of the complex in EtOH/CHCl<sub>3</sub>/concHCl/3hours provides red crystals of C<sub>5</sub>H<sub>5</sub>Fe(CO)<sub>2</sub>Cl (from CHCl<sub>3</sub>/petroleum ether 85:15, decomposing >87°). This chloride is not ionic but gives the orange cation C<sub>5</sub>H<sub>5</sub>Fe(CO)<sub>2</sub><sup>+</sup> with AgNO<sub>3</sub>/HNO<sub>3</sub>, and yellow crystals of C<sub>5</sub>H<sub>5</sub>Fe(CO)<sub>2</sub>CN (from CHCl<sub>3</sub>, decomposing >120°) by reaction with NaCN/MeOH/24 hours in 46% yield. [Piper, Cotton and Wilkinson *J Inorg Nucl Chem* **1** 165 1955, Beilstein **16** IV 1826.] When (C<sub>5</sub>H<sub>5</sub>Fe)<sub>2</sub>(CO)<sub>4</sub> is treated with NaHg it provides the anion C<sub>5</sub>H<sub>5</sub>Fe(CO)<sub>2</sub><sup>-</sup> that can be alkylated with MeI to furnish C<sub>5</sub>H<sub>5</sub>Fe(CO)<sub>2</sub>Me and related alkyl derivatives. [Aktogu et al. *J Organomet Chem* **262** 49 1984, R.B. King *Organometallic Synthesis Vol 1*, p 45, Academic Press, NY, 1965.]

**2,4-Cyclopentadien-1-yl lithium (lithium cyclopentadienide, CpLi)** [16733-97-4] **M 72.0**. CpLi is a white air and moisture sensitive solid which should be stored under N<sub>2</sub> preferably at low temperature. Like CpNa [4984-82-1] and CpK, it is handled in a dry box or Schlenk equipment in an inert atmosphere (compare with CpTi below). To avoid excessive manipulation, the reagents are sometimes prepared *in situ* in the reactions without being isolated. The solid is prepared in Schlenk equipment under a N<sub>2</sub> atmosphere by adding dropwise redistilled cyclopentadiene (2.34ml, 28.5mmol) to a solution of EtLi (28.5mmol) in \*benzene (50ml) and stirred magnetically until precipitation of a voluminous white solid is complete. This is filtered, or centrifuged, off

under N<sub>2</sub>, washed with dry Et<sub>2</sub>O and dried *in vacuo*. It is sparingly soluble in Et<sub>2</sub>O but is soluble in THF. [Wagner & Ebel *Tetrahedron* **26** 5155 1970.] For the preparation of a standard solution (use drybox, or Schlenk equipment under N<sub>2</sub>) cyclopentadiene (~10.0mmol) in dry THF (7ml) is refluxed with excess of lithium metal (see Ford *J Am Chem Soc* **82** 2857 1970) for 4 hours, filtered through a sintered glass funnel and diluted to a total volume of 10ml with THF which provides a 1.00M solution for further use. [Ford *J Organomet Chem* **32** 27 1971.] Its UV in THF has no maxima above 210nm, and its IR has  $\nu_{\max}$  (solid in Nujol) at 3048 (m), 2906 (w), 1742 (m), 1629 (m), 1548 (w), 1513 (w), 1426 (m), 1364 (sw), 1304 (w), 1258 (w), 1166 (w), 1110 (w), 1003 (s), 935 (sw), 887 (sw) and 746 (ss) cm<sup>-1</sup>; the <sup>1</sup>H NMR [100MHz, shift relative to β-CH<sub>2</sub> of THF (1.767ppm of 2% TMS in THF) at 27°] has a peak at δ 5.59 (at 1.0M) or 5.70 (at 0.01M) [Ford *J Organomet Chem* **32** 27 1971; the <sup>13</sup>C NMR (22.6MHz, MeOCH<sub>2</sub>CH<sub>2</sub>OMe/\*C<sub>6</sub>D<sub>6</sub> 3:1, under N<sub>2</sub> at 30°, TMS) has a multiplet with δ centered at 103.59 (m, <sup>1</sup>J<sub>C-H</sub> 159.4 and 6.9Hz). [Fischer et al. *J Organomet Chem* **116** 65 1976.]

**Cyclopentadienyl thallium [(η<sup>5</sup>-2,4-cyclopentadien-1-yl)thallium, thallium(I) cyclopentadienide, CpTI] [34822-90-7] M 269.5, blackens at ~60° in air and at ~230° in a sealed capillary, but does not melt below 270°.** CpTI is purified by repeated sublimation through a glass plug at ~80°/0.005mm, 75°/0.1mm or 100-110°/10mm to give pale yellow acicular (small pointed needles) crystals. It is moderately stable in air but darkens superficially after several months. As it is somewhat light sensitive, it should be stored in dark brown screw cap containers, and thus keeps for prolonged periods of time. It reacts with some solvents, e.g. CHCl<sub>3</sub> and CCl<sub>4</sub> (forming TlCl), and CS<sub>2</sub> (forming black and red materials), or is very poorly soluble, e.g. in THF, Et<sub>2</sub>O and \*C<sub>6</sub>H<sub>6</sub>. However, it is soluble enough in the latter solvents, even at low temperatures (e.g. -20°), to react almost completely to give high yields of products. *It is a more convenient reagent for cyclopentadienylation than CpNa, CpK or CpLi because it is less basic, can be prepared in aqueous medium, can be prepared on a large scale, and is stable in air at room temperature for long periods.* The cyclopentadiene ring of CpTI has IR in accord with C<sub>5v</sub> symmetry, exhibiting one C—H stretching band at  $\nu_{\max}$  3101 cm<sup>-1</sup> in the gas phase at 150° to 249° which compares with ferrocene (3109 cm<sup>-1</sup>), nickelocene (3107 cm<sup>-1</sup>), di-cyclopentadienylmanganese (3101 cm<sup>-1</sup>), di-cyclopentadienyl-magnesium(II) (3095 cm<sup>-1</sup>), and benzene (3099 cm<sup>-1</sup>), and is supported by theoretical considerations. The five membered ring is flat and ionic with the positive charge delocalised within it, and the Tl bears the counter negative charge and is centered on top of it. [Cotton & Raynolds *J Am Chem Soc* **80** 269 1958, Roberts et al. *J Mol Spectr* **35** 476 1970, Fritz *Chem Ber* **92** 780 1959.]

CpTI is conveniently prepared by dissolving or suspending a thallium(I) salt [TlCl 3.6g, TlBr 5.26g, TlI does not react, TlSCN 3.94g, TlNO<sub>3</sub> 4.00g or Tl(I) acetylacetonate 4.55g] in a solution of KOH (10g) in H<sub>2</sub>O (100ml) in a Waring Blender (900ml jar capacity), adding cyclopentadiene (2.0ml, freshly distilled [542-92-7]) and stirring for 30 seconds. Longer stirring gives a fine suspension that is difficult to collect. The white product is filtered off, washed with EtOH (2 x 10ml), dried in a desiccator over anhydrous CaSO<sub>4</sub> to give CpTI in 91, 93, 0, 99, 97 and 87% yields respectively which is usually used without purification. The preparation can be scaled up to give ~50g of CpTI, and the preferred salt is TlBr. If desired, further purification can be performed by high vacuum sublimation as described above. [Hunt & Doyle *Inorg Nucl Chem Lett* **2** 283 1966, Cotton & Raynolds *J Am Chem Soc* **80** 269 1958, Corey et al. *J Am Chem Soc* **93** 1489 1971.] **Note that all thallium compounds are POISONOUS and due precautions have to be exercised.**

CpTI is a source of the C<sub>5</sub>H<sub>5</sub> ligand, and reactions are performed under N<sub>2</sub> with highly purified solvents. The following are a few of the many reported reactions of CpTI. It has been used to prepare ferrocene, cobaltocene, nickelocene, also complexes such as Cp<sub>2</sub>TiCl<sub>2</sub> with TiCl<sub>4</sub> in \*C<sub>6</sub>H<sub>6</sub>, or of Cp(MeOCp<sub>2</sub>)Pd(II) from [CpPd(OMe)Cl]<sub>2</sub> with CpTI in \*C<sub>6</sub>H<sub>6</sub> [Hunt & Doyle *Inorg Nucl Chem Lett* **2** 283 1966], as well as compounds like the 1-methoxymethylcyclopenta-2,4-diene key intermediate for the synthesis of prostaglandins [Corey et al. *J Am Chem Soc* **93** 1489 1971]; it reacts with [CpMo(NO)I<sub>2</sub>]<sub>2</sub> in THF to give Cp<sub>2</sub>Mo(NO)I where the iodine can be replaced [King *Inorg Chem* **7** 90 1968], CpTI reacted with electrophilic olefins, e.g. tetracyanoethylene, in MeCN or THF to form compounds of the type Tl<sup>+</sup>[Cp(CN)C(CN)<sub>2</sub>]<sup>-</sup> [Freeman & Sneddon *Inorg Chem* **19** 1125 1980], and the reaction of CpTI (10-80% molar excess) with 7-chloronorbornadiene in dry diglyme at 150°/3-4 hours gave, after filtration of Tl salts, the hydrocarbon *hexahydro-3,4,7-methenocyclopenta[a]pentalene* in 8-12% yield from a one-step synthesis (a hydrocarbon which would have required several synthetic steps to prepare) [Battiste & Timberlake *J Org Chem* **42** 176 1977]. [Meister *Angew Chem* **69** 533 1957, Beilstein **16** IV 1690.]

**6,6-Dibenzyl-14-crown-4 (lithium ionophore VI, 6,6-dibenzyl-1,4,8,11-tetra-oxa-cyclo-tetradecane)** [106868-21-7] **M 384.5, m 102-103°**. Dissolve it in  $\text{CHCl}_3$ , wash this with saturated aqueous  $\text{NaCl}$ , dry ( $\text{MgSO}_4$ ), evaporate and purify it by chromatography on silica gel and gradient elution with  $^*\text{C}_6\text{H}_6/\text{MeOH}$  followed by preparative reverse phase HPLC on an octadecyl silanised silica (ODS) column and eluting with  $\text{MeOH}$ . It can be recrystallised from  $\text{MeOH}$  (IR has  $\nu_{\text{max}}$  at  $1120\text{ cm}^{-1}$ , C-O-C in  $\text{KBr}$ ). [Kimura et al. *Anal Chem* **59** 2331 1987, see also Maruyama et al. *J Chem Soc Perkin Trans 1* 2069 1986 and Tsukube et al. *J Chem Soc Perkin Trans 1* 1033 1986.] It complexes selectively with Li ions.

**Di-*n*-butyltin oxide** [818-08-6] **M 248.9, m >300°**. The oxide is prepared by hydrolysis of di-*n*-butyltin dichloride with  $\text{KOH}$ . Hence wash it with a little aqueous  $\text{M KOH}$ , then  $\text{H}_2\text{O}$  and dry at  $\sim 80^\circ/10\text{mm}$  until the IR is free from OH bands. [Cummings *Aust J Chem* **18** 98 1965, *Beilstein* **4** I 588.]

**Dicarbonyl(cyclopentadienyl)Co (I)** [12078-25-0] **M 180.1, b 75°/22mm, b 139-140°(dec)/710mm**. Best distilled in an atmosphere of  $\text{CO}$  in a vacuum. The red brown liquid decomposes slightly on distillation even in a vacuum to liberate some  $\text{CO}$ . Operations should be performed in an efficient fume cupboard. It is soluble in organic solvents and stable in air but decomposes slowly in sunlight and rapidly under UV. [Piper et al. *J Inorg Nucl Chem* **1** 165 1955, *Beilstein* **16** IV 1827.] **TOXIC**.

**Dichloro(2,2':6', 2''-terpyridine)platinum(II) dihydrate** [151120-25-1] **M 535.3, decomposes at 240-260°**. The aqueous filtrate from the reaction between terpyridyl (2.3g) and potassium platinumchloride (4.0g) in  $\text{H}_2\text{O}$  at  $90^\circ$  for 6 hours, is evaporated, cooled, and treated with  $\text{HCl}$  whereby the red chloride separates as the *dihydrate*, whereas the black *trihydrate* slowly crystallises from a cold aqueous solution and is air-dried. It is converted to the *dihydrate* in a desiccator over  $\text{H}_2\text{SO}_4$ , by washing with  $\text{EtOH}$ , heating in  $\text{H}_2\text{O}$  (slowly), or by precipitating from a warm aqueous solution with hydrochloric acid. This salt is not decomposed by boiling  $\text{HCl}$ , and is insoluble in most organic solvents. [Morgan & Burstall *J Chem Soc* 1498 1934, Lippard *Acc Chem Res* **118** 211 1987, *Beilstein* **26** IV 260.]

**Diethyl aluminium chloride** [96-10-6] **M 120.6, m -75.5°, b 106.5-108°/24.5mm,  $d_4^{20}$  0.96**. Distil it from excess dry  $\text{NaCl}$  (to remove ethyl aluminium dichloride) in a 50-cm column containing a heated nichrome spiral. [*Beilstein* **4** IV 4403.]

***N,N'*-Diheptyl-*N,N'*-5,5-tetramethyl-3,7-dioxanonanedi-*amide* [lithium ionophore I (ETH 149)]** [58821-96-8] **M 442.7**. Purify it by chromatography on Kieselgel using  $\text{CHCl}_3$  as eluent (IR has  $\nu_{\text{max}}$  at  $1640\text{cm}^{-1}$ ). [Kirsch et al. *Helv Chim Acta* **60** 2326 1977, Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Diphenylmercury** [587-85-9] **M 354.8, m 125.5-126°, 128-129°**. Sublime  $\text{Ph}_2\text{Hg}$ , then crystallise it from nitromethane or ethanol. If phenylmercuric halides are present, they can be converted to phenylmercuric hydroxide which, being much more soluble, remain in the alcohol or  $^*\text{benzene}$  used for crystallisation. Thus, crude material (10g) is dissolved in warm ethanol (*ca* 150ml) and shaken with moist  $\text{Ag}_2\text{O}$  (*ca* 10g) for 30 minutes, then heated under reflux for 30 minutes and filtered hot. Concentrating the filtrate by evaporation gives diphenylmercury, which is then recrystallised from  $^*\text{benzene}$  [Blair et al. *J Chem Soc* 3174 1959]. [*Beilstein* **16** IV 1702.] **TOXIC**.

**Disodium calcium ethylenediaminetetraacetate** [39208-14-5] **M 374.3**, (see  $\text{pKs}$  for EDTA in entry below). Dissolve it in a small amount of water, filter it and precipitate it with excess  $\text{EtOH}$ . Dry it at  $80^\circ$ . [*Beilstein* **4** IV 2451.]

**Disodium dihydrogen ethylenediaminetetraacetic acid ( $2\text{H}_2\text{O}$ )** [6381-92-6] **M 372.2, m 248°(dec),  $\text{pK}_1^{25}$  0.26,  $\text{pK}_2^{25}$  0.96,  $\text{pK}_3^{25}$  2.60,  $\text{pK}_4^{25}$  2.67,  $\text{pK}_5^{25}$  6.16,  $\text{pK}_6^{25}$  10.26 (see EDTA)**. Analytical reagent grade material can be used as primary standard after drying at  $80^\circ$ . Commercial grade material can be purified by crystallisation from water or by preparing a 10% aqueous solution at room temperature, then adding ethanol slowly until a slight permanent precipitate is formed, filtering, and adding an equal volume of ethanol. The precipitate is filtered off onto a sintered-glass funnel, is washed with acetone, followed by diethyl ether, and

dried in air overnight to give the *dihydrate*. Drying at 80° for at least 24 hours converts it to the *anhydrous* form. [Beilstein 4 IV 2451.]

**Disodium magnesium ethylenediaminetetraacetate** [14402-88-1] M 358.5,  $\text{pK}_1^{25}$  0.26  $\text{pK}_2^{25}$  0.96,  $\text{pK}_3^{25}$  2.60,  $\text{pK}_4^{25}$  2.67,  $\text{pK}_5^{25}$  6.16,  $\text{pK}_6^{25}$  10.26 (see EDTA). Dissolve it in a small amount of water, filter and precipitate it with an excess of MeOH. Dry it at 80°. [Beilstein 4 IV 2450.]

**Disodium succinate** [150-90-3] M 162.1. Crystallise it twice from water (1.2ml/g) and dry it at 125°. It has been freed from other metal ions by passage of a 0.1M solution through a column of Dowex resin A-1 (Na form). It is *hygroscopic*. [Beilstein 2 H 606, 2 IV 1908.]

**Di-*p*-tolylmercury** [50696-65-6, 537-64-4] M 382.8, m 244-246°. Crystallise it from xylene. [Whitmore et al. *Org Synth Coll Vol I* 231 1941, *Beilstein* 16 H 947, 16 I 667, 16 III 1329, 16 IV 1705.]

**Di-*p*-tolyl phenylphosphonate** [94548-75-1] M 388.3,  $n_D^{25}$  1.5758. Purify as described under diisooctyl phenylphosphonate.

**Eosin B (Bluish, Eosin Scarlet, 4',5'-dibromo-2',7'-dinitrofluorescein disodium salt)** [548-24-3] M 624.1,  $\lambda_{\text{max}}$  514nm, CI 45400. Free it from inorganic halides by repeated crystallisation from butan-1-ol. [Beilstein 19/6 V 469.]

**Eosin Y (as di-Na salt) (2',4',5',7'-tetrabromofluorescein di-Na salt)** [17372-87-1] M 691.9. Dissolve it in water and precipitate it by adding dilute HCl. The precipitate is washed with water, crystallised from ethanol, then dissolved in the calculated amount of dilute NaOH solution and evaporated to dryness on a water-bath. The purified disodium salt is then crystallised twice from ethanol [Parker & Hatchard *Trans Faraday Soc* 57 1894 1961]. [Beilstein 19 III/IV 2917.] [Same as below.]

**Ethylmercuric chloride** [107-27-7] M 265.1, m 192.5°, 193-194°. Mercuric chloride can be removed by suspending ethylmercuric chloride in hot distilled water, filtering with suction onto a sintered-glass crucible and drying it. Then crystallise it from ethanol and sublime it under reduced pressure. It can also be crystallised from water. [Marvel et al. *J Am Chem Soc* 47 3009 1925.]

**Ethylmercuric iodide** [2440-42-8] M 356.6, m 182°, 186°. Crystallise it once from water (50ml/g). [See previous entry, Marvel et al. *J Am Chem Soc* 4 3009 1925.]

**Ethylzinc (zinc diethyl)** [557-20-0] M 123.5, m -28°, -33.8°, b 116.8°/761mm,  $d_4^{25}$  1.205,  $n_D^{20}$  1.498. The presence of EtI in the liquid can be detected by its absorption spectrum. This can be removed by 2 or 3 passages over a Zn/Cu couple at ~150°. Any hydrocarbon impurities are removed by distillation at ~760mm in an inert atmosphere as it is flammable. It is moisture sensitive, hydrolysing to give ethane so store it in sealed ampoules under N<sub>2</sub> or argon.

It is also available as a ~15% solution in toluene ( $d_4^{20}$  0.915) or ~1M *n*-hexane ( $d_4^{20}$  0.726). Solutions should be stored under dry N<sub>2</sub> or argon, and if they contain some precipitate, the solutions should be filtered through a sintered frit under inert gas pressure. *Alternatively*, siphon carefully the required volume of supernatant. [Noller *Org Synth Coll Vol II* 184 1943, Bamford et al. *J Chem Soc* 471 1946, *Beilstein* 4 H 672, 4 I 609, 4 II 1044, 4 III 1999, 4 IV 4423.]

**Ethynyl tributylstannane** [994-89-8] M 315.1, b 76°/0.2mm, 130-135°/0.7mm, 200°/2mm,  $d_4^{20}$  1.1113,  $n_D^{20}$  1.4770. Purify the stannane by dissolving the reagent (*ca* 50g) in heptane (250ml), washing it with H<sub>2</sub>O (100ml), drying (MgSO<sub>4</sub>), evaporating and distilling in a vacuum. It has IR with  $\nu_{\text{max}}$  at 3280 (≡C-H), 2950, 2850, 2005 (C≡C), 1455, 1065 and 865cm<sup>-1</sup>. [Bottaro et al. *J Org Chem* 46 5221 1981, Stille & Simpson *J Am Chem Soc* 109 2138 1987, Zavgorodnii et al. *J Gen Chem USSR (Engl Edn)* 37 1469 1967.]

**Europium (III) acetate dihydrate** [62667-64-5] **M 383.1, pK<sub>1</sub><sup>25</sup> 8.31 (for aquo Eu<sup>3+</sup>)**. Recrystallise it several times from water [Ganapathy et al. *J Am Chem Soc* **108** 3159 1986]. [Beilstein **2** II 119.] For europium shift reagents see lanthanide shift reagents in "Aliphatic Compounds", Chapter 4.

**Ferrocene** [102-54-5] **M 186.0, m 173-174°**. Purify ferrocene by crystallisation from pentane or cyclohexane (also \*C<sub>6</sub>H<sub>6</sub> or MeOH can be used). It is moderately soluble in Et<sub>2</sub>O and sublimes readily above 100°. Crystallisation from EtOH gave material **m 172.5-173°**. [Wilkinson *Org Synth Coll Vol IV* 473 1963, Miller *J Chem Soc* 632 1952.] It has also been crystallised from methanol and sublimed *in vacuo*. [Saltiel et al. *J Am Chem Soc* **109** 1209 1987, Beilstein **16** IV 1783.]

**Ferrocene carboxaldehyde** [12093-10-6] **M 214.1, m 117-120°, 118-120°, 121°, 124.5°**. The aldehyde forms red crystals from heptane/CH<sub>2</sub>Cl<sub>2</sub>, EtOH or petroleum ether and sublimes at 70°/1mm. The *cyanohydrin* has **m 104°** (from \*C<sub>6</sub>H<sub>6</sub>/EtOH). The *semicarbazone* has **m 217-219°(dec)** after recrystallisation from aqueous EtOH. The *oxime* provides two isomers from petroleum ether *viz* **m 96-99°** and **m 155°**. The *O-acetyloxime* has **m 80-81°** after recrystallisation from hexane [Lindsay & Hauser *J Org Chem* **22** 355 1957]. The *2,4-dinitrophenylhydrazone* has **m 248°(dec)**. [Beilstein **16** IV 1798, Graham et al. *J Am Chem Soc* **79** 3416 1957, Broadhead et al. *J Chem Soc* 650 1958.]

**Ferrocene carboxylic acid** [1271-42-7] **M 230.1, m 210°(dec), 225-230°(dec), pK<sub>20</sub> 4.4 (H<sub>2</sub>O), 6.29 (68% aqueous MeOH)**. The acid crystallises as yellow crystals from petroleum ether (**m 225-230°dec**), CHCl<sub>3</sub> (**m 208.5°dec**), toluene/petroleum ether (**m 195-205°dec**), or aqueous ethanol. [Matsue et al. *J Am Chem Soc* **107** 3411 1985.] The *acid chloride* **m 49°** crystallises from pentane, and has UV with λ<sub>max</sub> at 458nm [Lau & Hart *J Org Chem* **24** 280 1959]. The *methyl ester* crystallises from aqueous MeOH with **m 70-71°**. The *anhydride* has **m 143-145°** when recrystallised from petroleum ether [Acton & Silverstein *J Org Chem* **24** 1487 1959]. The *amide* has **m 168-170°** when crystallised from CHCl<sub>3</sub>/Et<sub>2</sub>O or **m 167-169°** when crystallised from \*C<sub>6</sub>H<sub>6</sub>/MeOH. [Reeves *Org Synth* **56** 28 1977, Arimoto & Haven *J Am Chem Soc* **77** 6295 1955, Benkeser et al. *J Am Chem Soc* **76** 4025 1954.] [Beilstein **16** IV 1807.]

**Ferrocene-1,1'-dicarboxylic acid** [1293-87-4] **M 274.1, m >250°(dec), >300°, pK<sub>1</sub><sup>25</sup> 3.9, pK<sub>2</sub><sup>25</sup> 5.3**. The dicarboxylic acid crystallises in orange-yellow crystals from AcOH and sublimes above 230°. The *monomethyl ester* has **m 147-149°** [Nesmeyanov & Reutov *Dokl Acad Nauk USSR* **115** 518 1957]. The *dimethyl ester* has **m 114-115°** [Woodward et al. *J Am Chem Soc* **74**, 3458 1953]. The *diacid chloride* has **m 92-93°** when recrystallised from petroleum ether. [Nesmeyanov & Reutov *Dokl Acad Nauk SSSR* **120** 1267 1958, Kazitsyna et al. *Dokl Acad Nauk SSSR* **127** 333 1959, Beilstein **16** IV 1811.]

**Ferrocene-1,1'-dimethanol** [1291-48-1] **M 246.1, m 107-108°**. The diol is obtained from the diacid by LiAlH<sub>4</sub> reduction and recrystallised from Et<sub>2</sub>O/petroleum ether. [Reinhart et al. *J Am Chem Soc* **82** 4111 1960, Beilstein **16** IV 1795.]

**1-(Ferrocenyl)ethanol (α-methylferrocenemethanol)** [1277-49-2] **M 230.1, m 73-75°, 76-77°, 76-79°**. This versatile reagent is obtained by reduction of acetylferrocene (22.8g, see [1271-55-2]) in dry Et<sub>2</sub>O (500ml) with LiAlH<sub>4</sub> (1.9g) solution in anhydrous Et<sub>2</sub>O and refluxing for 2 hours. Excess of hydride is destroyed with EtOAc, the mixture is treated with a solution of NH<sub>4</sub>Cl (28g) in H<sub>2</sub>O, stirred for 0.5 hours at 0°, filtered, the organic layer is washed twice with H<sub>2</sub>O, dried (MgSO<sub>4</sub>), and evaporated to dryness. The residue (20.5g, 89%) is recrystallised from Et<sub>2</sub>O/petroleum ether to give the *carbinol* as orange rods. It is free from a C=O band but has a strong broad OH band in the IR spectrum. The *acetyl derivative* (Ac<sub>2</sub>O/pyridine/0°/15 hours), after sublimation at 60°/0.2mm, has **m 67-68°** and an IR band at 5.78μ (s, C=O), but no band in the OH region. [Arimoto & Haven *J Am Chem Soc* **77** 6295, Beilstein **16** IV 1975.]

**Germanium tetraethoxide** [14165-55-0] M 252.8, m  $-72^{\circ}$ , b  $54.5^{\circ}/5\text{mm}$ ,  $71-72^{\circ}/11\text{mm}$ ,  $188-190^{\circ}/722\text{mm}$ ,  $d_{25}^{25}$  1.1288. Distil  $\text{Ge}(\text{OEt})_4$  through a 10cm Vigreux column under reduced pressure. Alternatively, distil it through a Fenske glass helices column fitted with a total condensation variable take-off stillhead. Fractionate it under reduced pressure using a reflux ratio of 10:1. [Johnson & Fritz *J Am Chem Soc* 75 718 1953, Bradley *J Chem Soc* 4916 1956, Beilstein 1 IV 1308.]

**Hexabutyl-distannane** {hexabutyl-ditin, bis[(tributyl)tin]} [813-19-4] M 580.4, b  $160-162^{\circ}/0.3\text{mm}$ ,  $d_4^{20}$  1.148,  $n_D^{20}$  1.512. Purify bis[(tributyl)tin] by distilling it in a vacuum and store it in the dark. [Shirai et al. *Yakugaku Zasshi* 90 59 1970, *Chem Abstr* 72 90593 1970, Beilstein 4 I 590.]

**Hexamethylditin** {hexamethyldistannane, bis[(trimethyl)tin]} [661-69-8] M 327.6, m  $23.5^{\circ}$ , b  $85-88^{\circ}/45\text{mm}$ ,  $182^{\circ}/756\text{mm}$ ,  $d_{25}^{25}$  1.57. Wash bis[(trimethyl)tin] with  $\text{H}_2\text{O}$  and extract with  $^*\text{C}_6\text{H}_6$ , dry by filtering through powdered  $\text{Na}_2\text{SO}_4$ , remove  $^*\text{C}_6\text{H}_6$  on a rotary evaporator and fractionally distil the oily residue under vacuum (b  $85-88^{\circ}/45\text{mm}$ ). It boils at ca  $182^{\circ}$  at atmospheric pressure, but it cannot be distilled in air because the hot vapours flash in the condenser. [Kraus & Session *J Am Chem Soc* 47 2361 1925, Morris & Selwood *J Am Chem Soc* 63 2509 1941, Pedley et al. *Trans Faraday Soc* 53 1612 1957, Beilstein 4 IV 4346.]

**Iron(II) acetylacetonate** [iron(II) bis(2,4-pentanedionato- $\kappa\text{O}2,\kappa\text{O}4$ ),  $\text{Fe}(2^+)(\text{acac})_2$ ] [14024-17-0] M 254.1, m  $175^{\circ}$ . The preparation and handling of this complex should be done in a dry box or Schlenk equipment under dry pre-purified  $\text{N}_2$ ; and solvents should be de-gassed prior to use.  $\text{Fe}(\text{acac})_2$  is prepared by adding  $\text{FeCl}_2 \cdot 4\text{H}_2\text{O}$  (25g) in degassed  $\text{H}_2\text{O}$  (50ml) containing a small amount of sodium dithionite to reduce any Fe(III) impurities, to a degassed aqueous solution (300ml) containing piperidine (27.3ml) and redistilled acetylacetonone (28.1ml, [123-54-6]). After 15 minutes the resulting precipitate is filtered off and washed with degassed  $\text{H}_2\text{O}$ , EtOH then  $\text{Et}_2\text{O}$  and the yellow-brown powder [presumed to be  $\text{Fe}(\text{acac})_2 \cdot 2\text{H}_2\text{O}$  Emmert & Jarczyński *Chem Ber* 64 1072 1931] is transferred into a sublimation apparatus with a large cold finger and pumped dry at  $10^{-3}$  mm for 6 hours; and the hydrated complex is then heated at  $90^{\circ}/10^{-3}\text{mm}$  for at least 12 hours to dehydrate it and the orange-brown solid is sublimed twice at  $165-175^{\circ}/10^{-3}\text{mm}$ . The small orange-brown crystals that sublime at first become darker as the crystals grow, giving finally larger black clusters which are analytically pure anhydrous solid. Note that the complex is very sensitive to  $\text{O}_2$  in air, darkens in colour upon oxidation, and great care should be exercised to avoid oxidation when using it. It maybe *pyrophoric* when dry. Its molecular weight (cryoscopic and ebullioscopic) in  $^*\text{C}_6\text{H}_6$  increases as the concentration increases reaching a maximum for a hexamer, and the data are supported by the UV absorption spectra. [Buckingham et al. *Aust J Chem* 20 281 1967, Dwyer & Sargeson *Pr J Soc NSWales* 90, 141, 142 1956; for the formation constant see Izatt et al. *J Phys Chem* 59 80 1955, Beilstein 1 III 3122, 1 IV 3675.] Note: iron(1+) bis(2,4-pentanedionato- $\kappa\text{O}2,\kappa\text{O}4$ ) has [20149-10-4], see Chapter 6, Catalysts-Part 1 for  $\text{Fe}(\text{III})(\text{acac})_3$ , [14024-18-1].

**Iron(II) bis(1,1,1,5,5,5-hexafluoropentan-2,4-dionate)** [iron(II) bis(1,1,1,5,5,5-hexafluoro-2,4-pentanedionato- $\kappa\text{O}2,\kappa\text{O}4$ ),  $\text{Fe}(2^+)(\text{facac})_2$ ,  $\text{Fe}(\text{hfacac})_2$ ] [28736-68-7] M 470.1, mp  $40^{\circ}$ . The preparation and handling of this complex [cf  $\text{Fe}(\text{acac})_2$  14024-17-0] should be carried out in a dry box or Schlenk equipment under dry pre-purified  $\text{N}_2$ ; and solvents should be de-gassed prior to use.  $\text{Fe}(\text{hfacac})_2$  is prepared by adding  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  (11.8g) in  $\text{H}_2\text{O}$  (50ml) slowly to a mixture of redistilled hexafluoroacetylacetonone (12.2ml, see [1522-22-1]) and piperidine (8.5ml), whereby a dark violet precipitate separates immediately. This solid is filtered off, washed with a large volume of  $\text{H}_2\text{O}$  (1000 ml), placed in a sublimation apparatus and evacuated at  $10^{-3}\text{mm}$  at  $\sim 25^{\circ}$  for hours. The temperature is slowly raised to  $40^{\circ}$  and kept at  $45-50^{\circ}$  and  $10^{-3}\text{mm}$  for 3 days, and the dark violet crystals that formed on the cold finger are collected, and if necessary are ground to a fine powder and re-sublimed to give analytically pure  $\text{Fe}(\text{hfacac})_2$  (IR is free of bands at  $3300-3500\text{ cm}^{-3}$ , i.e. absence of  $\text{H}_2\text{O}$ ). It requires a much longer period of evacuation than  $\text{Fe}(\text{acac})_2$  for dehydration. The strong electron withdrawing effects of the  $\text{CF}_3$  groups make this complex relatively more stable to aerial oxidation than  $\text{Fe}(\text{acac})_2$ . It can be recrystallised from very dry  $^*\text{C}_6\text{H}_6$  without alteration, where its molecular weight (cryoscopic and ebullioscopic) increases with increase in concentration up to a value of 1.7 (compare with 6 for  $\text{Fe}(\text{acac})_2$  above). [Buckingham et al. *Aust J Chem* 20 281 1967.]

**Iron(II) dicyanobis(2,2'-dipyridine) trihydrate**  $[\text{Fe}(\text{bpy})_2(\text{CN})_2 \cdot 3\text{H}_2\text{O}]$  [15603-10-8] **M 474.3, dec. on heating.** The iron complex is prepared from 2,2'-dipyridine (4.7g, 30mmol) and ferrous ammonium sulfate hexahydrate (3.9g, 10mmol) in  $\text{H}_2\text{O}$  (400ml) which are heated to just below boiling, and a freshly prepared solution of KCN (10g) in  $\text{H}_2\text{O}$  (20ml) is added all at once to it, stirred, and the hot mixture is allowed to cool to room temperature. The complex separates as very dark violet (almost black) fine crystals rapidly, but is allowed to stand for 1 hour. The solid is collected and washed liberally with  $\text{H}_2\text{O}$ . Unused bipyridine is recovered from the filtrate by extraction with  $\text{Et}_2\text{O}$ . The iron complex can be purified by dissolving it in concentrated  $\text{H}_2\text{SO}_4$  (25ml) and slowly diluting it with  $\text{H}_2\text{O}$  (800ml, CARE as it will warm up) with stirring. The crystalline solid is collected washed free from  $\text{H}_2\text{SO}_4$  with small volumes of  $\text{H}_2\text{O}$ , sucked dry, rinsed with  $\text{Me}_2\text{CO}$  and dried *in vacuo* to give analytically pure *trihydrated complex* (3.5g, 75%). It is soluble in  $\text{H}_2\text{O}$  to give a pale red solution, in  $\text{EtOH}$  to give a deep red-violet liquid, and in  $\text{CHCl}_3$  to give a deep blue colour; but is only very slightly soluble in basic solution. Large amounts can be dissolved in concentrated  $\text{H}_2\text{SO}_4$  to give the yellow di-protonated species which on gradual dilution with  $\text{H}_2\text{O}$  the solution changes its hue to orange then to red and finally to the dark violet colour of the original crystals. It is a diamagnetic Fe(II) complex which is readily oxidised to the corresponding Fe(III) complex and can be titrated with Ce(IV) sulfate in  $\text{H}_2\text{SO}_4$ . On titration with  $\text{HClO}_4$  in anhydrous  $\text{AcOH}$  (Care due to possible explosion; Schilt *J Am Chem Soc* **82** 5779 1960), the first equivalence point leads to a mono-protonated orange species, and at the second equivalence point (not quite as sharp as the first) it provides the yellow di-protonated species. The electronic (Schilt *J Am Chem Soc* **82** 3000 1960, Schilt *J Am Chem Soc* **82** 5779 1960, Madeja & König *J Inorg Nucl Chem* **25** 377 1963), and vibrational spectra (Schilt *Inorg Chem* **3** 1323 1964) have been reported, and the *cis*-configuration has been deduced from IR (Hamer & Orgel *Nature* **190** 439 1961, Schilt *Inorg Chem* **3** 1323 1964) and stereochemical (Madeja *Chem Zvesti* **19** 186 1965) considerations. [Schilt *Inorg Synth* **12** 249 1970.]

**Iron(III) ethoxide**  $[\text{Fe}(\text{OEt})_3]$  [5058-42-4] **M 91.0, m 120°, volatilities at 155°/0.1mm in a molecular still.** The ethoxide is prepared by adding excess of ammonia to a solution of anhydrous  $\text{FeCl}_3$  (15g) in  $^*\text{C}_6\text{H}_6$  (170ml) and  $\text{EtOH}$  (76ml), which results in an exothermic reaction at the end of which the mixture is evaporated to dryness *in vacuo*. The residue is extracted with  $^*\text{C}_6\text{H}_6$  (150ml), the  $\text{NH}_4\text{Cl}$  is filtered off, the filtrate is evaporated to dryness *in vacuo* to leave a viscous brown residue. This is dissolved in hot  $\text{EtOH}$  (30ml) and brown crystals of analytically pure *ferric ethoxide* (4.2g) deposit slowly. It is trimeric in  $^*\text{C}_6\text{H}_6$  solution (molecular weight by ebullioscopy) and is for preparing other useful ferric trialkoxides by alcohol exchange in benzene solution, e.g. with *n*- $\text{PrOH}$ , *n*- $\text{BuOH}$ , *iso*- $\text{BuOH}$  or *n*-amyl alcohol, via the azeotropic removal of  $\text{EtOH}$ . It is commercially available as a 1M solution in  $\text{EtOH}$ . [Bradley et al. *J Chem Soc* 126 1955, *Beilstein* **1** IV 1243.]

**Iron(III) meso-5,10,15,20-tetraphenylporphine chloride complex** [5,10,15,20-tetraphenyl-21H, 23H-porphine iron(III) chloride] [16456-81-8] **M 704.0,  $\lambda_{\text{max}}$  418nm.** Purify the complex by extraction from a thimble (Soxhlet) with  $\text{CHCl}_3$ . Concentrate the extract to *ca* 10ml and add *ca* 80ml of hot  $\text{MeOH}$ . Dark blue crystals separate on cooling. It can be recrystallised several times from  $\text{CHCl}_3/\text{MeOH}$ . Avoid prolonged heating. It is quite soluble in organic solvents but insoluble in petroleum ether. [Rothemund & Manotti *J Am Chem Soc* **70** 1808 1948, UV: Dorrough et al. *J Am Chem Soc* **73** 4315 1951, *Beilstein* **26** III 1960.] [As catalyst for silylation of OH groups see Firouzabadi et al. *Synth Commun* **27** 2709 1997.]

**Lanthanide shift reagents** See in "Aliphatic Compounds", Chapter 4, europium (III) acetate above and  $\text{Eu}(\text{tmc})_3$  and  $\text{Eu}(\text{tfc})_3$  below.

**Lead II acetate (sugar of lead)** [301-04-2 (anhydrous), 6080-56-4 ( $3\text{H}_2\text{O}$ )] **M 325.3, m 280°,  $\text{pK}_1^{25}$  7.1 (for  $\text{Pb}^{2+}$ ),  $\text{pK}_2^{25}$  10.1 ( $\text{HPbO}_2^-$ ),  $\text{pK}_3^{25}$  10.8 ( $\text{PbO}_2^{2-}$ ).** Crystallise  $\text{Pb}(\text{OAc})_2$  twice from anhydrous acetic acid and dry it under vacuum for 24 hours at 100°. The solubility in  $\text{H}_2\text{O}$  is 63% (at ~20°) and 200% (at boiling point). [*Beilstein* **2** IV 118.]

**Lead (bis-cyclopentadienyl) (plumbocene)** [1294-74-2] **M 337.4.** Purify it by vacuum sublimation. Handle and store it under  $\text{N}_2$ . [Dave et al. *J Chem Soc* 3686 1979, *Beilstein* **16** IV 1614.]

**Lead tetraacetate** [546-67-6] **M 443.2, m 175-180°**. Colourless prisms or needles purified by dissolving in hot glacial acetic acid containing a little acetic anhydride, treated with decolorising charcoal, collected on a hot water funnel or preheated Büchner funnel with minimum contact with moist air, and dried in a vacuum desiccator over KOH pellets. Store it in a well-stoppered vessel as it is decomposed by moisture to form brown PbO<sub>2</sub>. It attacks skin, is soluble in hot AcOH, \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub>, tetrachloroethane and is used as a powerful oxidising agent. [Bailar *Inorg Synth* **1** 42 1939, Baudler in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **I** p 767 1963, *Beilstein* **2** IV 118.] **TOXIC**.

**Lithium benzoate** [553-54-8] **M 128.1**. Crystallise the salt from EtOH (13ml/g) by partial evaporation. [*Beilstein* **9** I 107, **9** IV 279.]

**Lithium diisopropylamide** [4111-54-0] **M 107.1, b 82-84°/atm, 84°/atm, d<sup>22</sup> 0.722, flash point -6°**. LiN(*iso*-Pr)<sub>2</sub> is purified by refluxing over Na wire or NaH for 30 minutes and then distilled into a receiver under N<sub>2</sub>. Because of the low boiling point of the amide, a dispersion of NaH in mineral oil can also be used directly in this purification without prior removal of the oil. It is **HIGHLY FLAMMABLE**, and is decomposed by air and moisture. [Wittig & Hesse *Org Synth* **50** 69 1970, *Beilstein* **4** H 154, **4** I 369, **4** II 630, **4** III 274, **4** IV 510.]

**Lithium formate monohydrate** [6108-23-2 (H<sub>2</sub>O), 556-63-8 (anhydrous)] **M 70.0, d<sub>4</sub><sup>20</sup> 1.46**. Crystallise HCO<sub>2</sub>Li from hot water (0.5ml/g) by chilling. [*Beilstein* **2** III 22, **2** IV 13.]

**Lithium methylate (lithium methoxide)** [865-34-9] **M 38.0**. The most probable impurity is LiOH due to hydrolysis by moisture. It is important to keep the sample dry. It can be dried by keeping in a vacuum at 60-80° under dry N<sub>2</sub> using an oil pump for a few hours. Store it under N<sub>2</sub> in the cold. It should not have bands above 3000cm<sup>-1</sup>, and its IR (KBr) should have  $\nu_{\max}$  at 1078, 2790, 2840 and 2930cm<sup>-1</sup>. [Suebold *J Org Chem* **21** 156 1956, *Beilstein* **1** IV 1220, 1241.]

**Lithium pentamethylcyclopentadienide (LiCp', 1,2,3,4,5-pentamethyl-2,4-pentadiene-1-ylithium)** [51905-34-1] **M 142.2, m >230°**. In the necessary Schlenk equipment under argon or dry N<sub>2</sub> an equimolar amount of 4M *n*-BuLi in hexane is added dropwise to a stirred solution of equimolar pentamethylcyclopentadiene (Cp') in dry THF (30ml per g of Cp') at 0° which gives a pulpy pale yellow suspension. Magnetic stirring is continued for 2 hours while the temperature is allowed to rise to ~25°; and this can be used directly for coordinating a Cp' ligand to metals. If LiCp' is to be isolated, then the solvent to be used is petroleum ether (~15ml per g of Cp'), addition of the *n*-BuLi solution is done within a few minutes without cooling as the reaction is slower than in THF, and stirring is continued for 24 hours. The solvent is removed *in vacuo*, or the colourless to pale grey solid (>90% yield) is filtered off, washed with petroleum ether under N<sub>2</sub> and dried *in vacuo*. [Kohl et al. in *Organometallic Synthesis* (R. Bruce King and J.J. Eisch) **vol 3** 381, Elsevier Amsterdam 1986]. LiCp' is air and moisture sensitive, and should be stored and weighed in an inert atmosphere. [King & Bisnette *J Organomet Chem* **8** 287 1967.] This reagent has been prepared also *in situ* from Cp' with MeLi/THF at 0° [White et al. *Synth Commun* **3** 425 1973], with *n*-BuLi in dimethoxyethane [Manriques & Bercaw *J Am Chem Soc* **96** 6229 1974], or as a slurry of LiCp' in Et<sub>2</sub>O at ~25° [Beachley et al. *Organometallics* **4** 1675 1985]. The metal-Cp' bond in LiCp' has a more covalent nature than that in KCp' and NaCp' which are more ionic [Jutzi et al. *Chem Ber* **118** 1959 1985].

**Lithium picrate** [18390-55-1] **M 221.0**. Recrystallise the picrate three times from EtOH and dry it under vacuum at 45° for 48 hours [D'Aprano & Sesta *J Phys Chem* **91** 2415 1987]. [*Beilstein* **6** H 276, **6** II 263, **6** III 880, **6** IV 1390.] The necessary precautions should be taken in case of **EXPLOSION**.

**Lithium salicylate** [552-38-5] **M 144.1**. Recrystallise the salicylate from EtOH (2ml/g) by partial evaporation. [*Beilstein* **10** H 59, **10** II 32, **10** III 93, **10** IV 126.]

**Magnesium acetate** [142-72-3 (anhydrous), 16674-78-5 (4H<sub>2</sub>O)] **M 214.5, m 80°**. Crystallise it from anhydrous acetic acid, then dry it under vacuum for 24 hours at 100°. [Nencollas *J Chem Soc* 744 1956,

*Beilstein 2* IV 113.]

**Magnesium benzoate trihydrate** [553-70-8] **M 320.6, m ~200°**. Crystallise it from water (6ml/g) between 100° and 0°. [*Beilstein 9* III 376, **9** IV 280.]

**Magnesium ethylate (magnesium ethoxide)** [2414-98-4] **M 114.4**. Dissolve *ca* 1g of solid in 12.8ml of absolute EtOH and 20ml of dry xylene, and reflux in a dry atmosphere (use CaCl<sub>2</sub> in a drying tube at the top of the condenser). Add 10ml of absolute EtOH and cool. Filter the solid under dry N<sub>2</sub> and dry it in a vacuum. *Alternatively*, dissolve it in absolute EtOH and pass it through molecular sieves (40 mesh) under N<sub>2</sub>, evaporate under N<sub>2</sub>, and store it in a tightly stoppered container. [Smith & Wiley *J Am Chem Soc* **68** 889 1946, *Beilstein 1* III 1283.]

**Magnesium D-gluconate** [3632-91-5] **M 414.6, [α]<sub>D</sub><sup>20</sup> +13.5°, [α]<sub>D</sub><sup>20</sup> +11.3° (c 1, H<sub>2</sub>O)**. Crystallise it from dilute EtOH to give *ca trihydrate*, and then dry it at 98° in high vacuum. It is insoluble in EtOH, and the solubility in H<sub>2</sub>O is 16% at 25°. [Prescott et al. *Ind Eng Chem* **45** 338 1953, *Beilstein 3* IV 1256.]

**Magnesium ionophore I (ETH 1117), (N,N'-diheptyl-N,N'-dimethyl-1,4-butanediamide)** [75513-72-3] **M 340.6**. Purify it by flash chromatography (at 40 kPa) on silica and eluting with EtOH/hexane (4:1). Its IR (CHCl<sub>3</sub>) has  $\nu_{\max}$  at 1630 cm<sup>-1</sup>. [Eene et al. *Helv Chim Acta* **63** 2271 1980.] It is a good magnesium selectophore compared with Na, K and Ca [Lanter et al. *Anal Chem* **52** 2400 1980].

**Magnesium ionophore II (ETH 5214), [N,N]-octamethylene-bis(N'-heptyl-N-methyl methylmalonamide)** [119110-37-1] **M 538.8**. The reagent (*ca* 700mg) can be purified by flash chromatography on Silica Gel 60 (30g) and eluting with CH<sub>2</sub>Cl<sub>2</sub>/Me<sub>2</sub>CO (4:1). [Hu et al. *Anal Chem* **61** 574 1989.]

**Magnesium lactate** [18917-37-1] **M 113.4**. Recrystallise the salt from water (6ml/g) between 100° to 0°. [*Beilstein 3* IV 636.]

**Magnesium succinate** [556-32-1] **M 141.4**. Recrystallise the salt from water (0.5ml/g) between 100° and 0°. [*Beilstein 2* IV 1912.]

**Magon [3-hydroxy-4-(hydroxyphenylazo)-2-naphthoyl-2,4-dimethylanilide, Xylidyl Blue II]** [523-67-1] **M 411.5, m 246-247°**. Suspend it in H<sub>2</sub>O and add aqueous NaOH until it dissolves, filter and acidify with dilute HCl. Collect the dye, dissolve it in hot EtOH (solubility is 100mg/L at *ca* 25°) concentrate to a small volume and allow to cool. The solubility of the Na salt in H<sub>2</sub>O is 0.4mg/ml. [Mann & Yoe *Anal Chim Acta* **16** 155 1957, Mann & Yoe *Anal Chem* **28** 202 1956.]

**Manganese (II) acetylacetonate** [14024-58-9] **M 253.2, m ~250°, 261°(dec)**. Purify it by stirring 16g of reagent for a few minutes with 100ml absolute EtOH and filter by suction as rapidly as possible through coarse filter paper. Sufficient EtOH is added to the filtrate, to make up for the loss of EtOH and to redissolve any solid that separates. Water (15ml) is added to the filtrate, and the solution is evaporated with a stream of N<sub>2</sub> until reduced to half its volume. Cool for a few minutes and filter off the yellow crystals, dry them under a stream of N<sub>2</sub>, then in a vacuum at room temperature for 6-8 hours. These conditions are important for obtaining the *dihydrate*. A vacuum to several mm of Hg or much lower pressure for several days produces the *anhydrous complex*. The degree of hydration can be established by determining the loss in weight of 100g of sample after heating for 4 hours at 100° and <20mmHg. The theoretical loss in weight for 2H<sub>2</sub>O is 12.5%. The material sublimes at 200°/2mm. It is soluble in heptane, MeOH, EtOH or \*C<sub>6</sub>H<sub>6</sub> at 30°. [Charles *Inorg Synth* **VI** 164 1960, Fernelius & Biswas *Inorg Synth* **V** 105 1957, *Beilstein 3* II 3122.]

**Manganous acetate tetrahydrate** [6156-78-1 (4H<sub>2</sub>O), 638-38-0 (anhydrous)] **M 245.1, m 80°, d<sub>4</sub><sup>20</sup> 1.59, pK<sup>25</sup> 10.59 (for Mn<sup>2+</sup> hydrolysis)**. Crystallise it from water acidified with acetic acid. [*Beilstein 2* IV 120.]

**Manganous lactate trihydrate** [6505-50-6 3 H<sub>2</sub>O, 51877-53-3 (xH<sub>2</sub>O)] **M 287.1**. Recrystallise the lactate

from water. [IR: Ranade & Biswas *J Indian Chem Soc* **44** 314 1967, *Beilstein* **3** IV 637.]

**Mercuric(II) acetate** [ $\text{Hg}(\text{OAc})_2$ ] [1600-27-7] **M 318.7, m 178-180°, 179-182°, pK<sub>1</sub><sup>25</sup> 2.47, pK<sub>2</sub><sup>25</sup> 3.49 (for Hg<sup>2+</sup> hydrolysis).** It is prepared from HgO and acetic acid. Recrystallise it from glacial acetic acid. Its solubility in H<sub>2</sub>O is 1g/2.5ml at ~20° and 1g/ml at 100°, and it is soluble in EtOH. The solid is light sensitive and is hydrolysed slowly in H<sub>2</sub>O to yield a yellow oxide. Store it in a tightly stoppered bottle away from light. It is a reagent for mercuration of alkenes (and aromatic compounds, cf mercury trifluoroacetate below) [Larock *Tetrahedron* **38** 1713 1928]; de-mercuration being achieved with NaBH<sub>4</sub>. [Wagenknecht & Juza in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol **II** p 1120-1121 1965, *Beilstein* **2** IV 114.] **POISONOUS.**

**Mercury(II) bis(cyclopentadienyl)** [18263-08-6] **M 330.8 starts to dec at ~60° but melts at 83-85°.** The pale yellow crystals can be purified by low-temperature (-78°) recrystallisation from Et<sub>2</sub>O. Store it at low temperature in the dark as it decomposes slowly at room temperature even in the dark. It reacts with FeCl<sub>2</sub> in tetrahydrofuran to give ferrocene quantitatively. It is soluble in Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub>, CCl<sub>4</sub>, and most organic solvents. [Campbell et al. *J Am Chem Soc* **94** 8387 1972, Wilkinson & Piper *J Inorg Nucl Chem* **2** 32, 36 1956, Piper & Wilkinson *J Inorg Nucl Chem* **3** 104,106 1956, *Beilstein* **16** IV 1702.]

**Mercury dibromofluorescein {mercurochrome, merobromin, [2',7'-dibromo-4'-(hydroxy-mercurio)-fluorescein di-Na salt]}** [129-16-8] **M 804.8, m>300°.** The Na salt is dissolved in the minimum volume of H<sub>2</sub>O, or the free acid suspended in H<sub>2</sub>O and dilute NaOH is added to cause it to dissolve, filter and acidify it with dilute HCl. Collect the precipitate, wash it with H<sub>2</sub>O by centrifugation and dry it in a vacuum. The di Na salt can be purified by dissolving it in the minimum volume of H<sub>2</sub>O and is precipitated by adding EtOH, filter, wash it with EtOH or Me<sub>2</sub>CO and dry it in a vacuum. Its solubility in 95% EtOH is 2% and in MeOH it is 16%. [White *J Am Chem Soc* **42** 2355 1920.]

**Mercury Orange [1-(4-chloromercuriophenylazo)-2-naphthol]** [3076-91-3] **M 483.3, m 291.5-293°(corr) with bleaching.** Wash it several times with boiling 50% EtOH and recrystallise it from 1-butanol (0.9g/l of boiling alcohol). The fine needles are insoluble in H<sub>2</sub>O but slightly soluble in cold alcohols, CHCl<sub>3</sub> and soluble in aqueous alkalis. [Chaikin *J Am Chem Soc* **70** 3522 1948.]

**Mercury(II) trifluoroacetate** [13257-51-7] **M 426.6, m 171-173°.** It is prepared from HgO and trifluoroacetic acid and is recrystallised from trifluoroacetic anhydride/trifluoroacetic acid. It is a useful reagent for mercuration of aromatic compounds [Lau & Kochi *J Am Chem Soc* **108** 6720 1986]. Store it in a tightly stoppered bottle away from light. [cf Aylett in *Comprehensive Inorganic Chemistry* (Bailer et al. eds) Vol **3** p187 1973, *Beilstein* **2** IV 458.] It is very **TOXIC** and *hygroscopic*.

**Methylmercuric chloride** [115-09-3] **M 251.1, m 167°, 170°.** Recrystallise it from absolute EtOH (20ml/g). Its UV has  $\lambda_{\text{max}}$  at 206nm ( $\epsilon$  1.37). [See EtHgCl above; Breitingner et al. *J Organomet Chem* **256** 217 1983, Slotta et al. *J Prakt Chem* **120** 249 1929, Waugh et al. *J Phys Chem* **59** 395 1955, *Beilstein* **16** IV 1729.]

**Milling Yellow G** [51569-18-7]. Salted out three times with sodium acetate, then repeatedly extracted with EtOH. [McGrew & Schneider *J Am Chem Soc* **72** 2547 1950, *Beilstein* **16** II 125.] See Solochrome Violet R [2092-55] in "Aromatic compounds", Chapter 4.

**Monoperoxyphthalic acid magnesium hexahydrate salt. (MMPP)** [84665-66-7] **M 494.7, m ~93°(dec).** MMPP is a safer reagent than *m*-chloroperbenzoic acid for oxidation reactions because it is not as explosive and has advantages of solubility. It is soluble in H<sub>2</sub>O, low-molecular-weight alcohols, *i*-PrOH and DMF. The product of reaction, Mg phthalate, is soluble in H<sub>2</sub>O. It has been used in aqueous phase to oxidise compounds in e.g. CHCl<sub>3</sub> and using a phase transfer catalyst, e.g. methyltrioctylammonium chloride [Brougham *Synthesis* 1015 1987]. The oxidising activity can be checked (as for perbenzoic acid in Silbert et al. *Org Synth Coll Vol V* 906 1973), and if found to be low it would be best to prepare afresh from phthalic anhydride (1mol), H<sub>2</sub>O<sub>2</sub> (1mol) and MgO at 20-25° to give MMPP. [Hignett, European Pat Appl 27 693 1981, *Chem Abstr* **95** 168810 1981, *Beilstein* **9** IV 3260.]

**Neopentoxo lithium** [3710-27-8] **M 94.1**. Recrystallise alkoxide from hexane [Kress & Osborn *J Am Chem Soc* **109** 3953 1987]. [*Beilstein* 1 H 406, 1 I 201, 1 II 435, 1 III 1448, 1 IV 1690, for alcohol and Al and Fe neopentoxides.]

**New Methylene Blue N (2,8-dimethyl-3,7-bis(ethylamino)phenothazinium chloride 0.5 ZnCl<sub>2</sub>)** [6586-05-6] **M 416.1**, **m >200°(dec)**, **pK<sub>1</sub> 3.54**, **pK<sub>2</sub> 4.82**. Crystallise the dye from \*benzene/MeOH (3:1). [*Beilstein* **27** IV 5163.]

**Nickel (II) acetate (4H<sub>2</sub>O)** [6018-89-9] **M 248.9**, **d<sub>4</sub><sup>20</sup> 1.744**, **pK<sub>1</sub><sup>25</sup> 8.94 (from Ni<sup>2+</sup> hydrolysis)**. The acetate crystallises from aqueous AcOH as the green *tetrahydrate salt*. It is soluble in 6 parts of H<sub>2</sub>O. It forms lower hydrates and should be kept in a well-sealed container. [Hardt & Pohlmann *Z Anorg Allgem Chem* **343** 92 1966, *Beilstein* **2** IV 120.]

**Nickelocene [bis-(cyclopentadienyl)nickel II]** [1271-28-9] **M 188.9**, **m 173-174°(under N<sub>2</sub>)**. Dissolve the complex in Et<sub>2</sub>O, filter and evaporate in a vacuum. Purify it rapidly by recrystallisation from petroleum ether using a solid CO<sub>2</sub>/Me<sub>2</sub>CO bath, and has **m 171-173°** (in an evacuated tube). Also purify it by vacuum sublimation. [Wilkinson et al. *J Am Chem Soc* **76** 1970 1954, *Beilstein* **16** IV 1831.]

**Nickel (II) phthalocyanine** [14055-02-8] **M 571.3**, **m >300°**. Wash it well with H<sub>2</sub>O and boiling EtOH and sublime it at high vacuum in a slow stream of CO<sub>2</sub>. A special apparatus is used (see reference), with the phthalocyanine being heated to red heat. The sublimate is in the form of needles with an extremely bright red lustre. The powder is dull greenish blue in colour. [Barrett et al. *J Chem Soc* 1719 1936, *Beilstein* **26** III/IV 4255.]

**Nickel 5,10,15,20-tetraphenylporphyrin** [14172-92-0] **M 671.4**, **λ<sub>max</sub> 414(525)nm**. Purify it by chromatography on neutral (Grade I) alumina, followed by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/MeOH [Yamashita *J Phys Chem* **91** 3055 1987]. [*Beilstein* **26** III/IV 1960.]

**Octadecyl isonicotinate** See hydrogen ionophore IV, ETH 1778 in "Heterocyclic Compounds" [103225-02-1], Chapter 4.

**Phenylmercuric acetate (PhHgOAc)** [62-38-4] **M 336.7**, **m 148-151°, 149°, 151.8-152.8°**. This acetate forms small colourless lustrous prisms from EtOH. Its solubility in H<sub>2</sub>O is 0.17%, but it is more soluble in EtOH, Me<sub>2</sub>CO and \*C<sub>6</sub>H<sub>6</sub>. [Maynard *J Am Chem Soc* **46** 1510 1925, Coleman et al. *J Am Chem Soc* **59** 2703 1937, *J Am Pharm Assoc* **25** 752 1936, *Beilstein* **16** IV 1720.] See PhHgOH below.

**Phenylmercuric hydroxide (PhHgOH)** [100-57-2] **M 294.7**, **m 195-203°**, **pK<sub>25</sub> 10 (~4.5)**. Crystallise it from dilute aqueous NaOH. [Waugh et al. *J Phys Chem* **59** 395 1955, *Beilstein* **16** IV 1721.]

**Phenylmercuric nitrate (PhHgNO<sub>3</sub>)** [8003-05-2] **M 634.4**, **m 178-188°**. When recrystallised from water or aqueous EtOH, it has **m 191-191.5°**, and from 95% EtOH it has **m 186.5°** [Challanger & Rothstein *J Chem Soc I* 1258 1934]. When recrystallised from \*C<sub>6</sub>H<sub>6</sub> or CHCl<sub>3</sub>, it has **m 131-131.5°** [Morton et al. *J Am Chem Soc* **69** 908 1947]. [*Beilstein* **16** III 1348, **16** IV 1721.]

**Phthalocyanine** [574-93-6] **M 514.6**. Purify phthalocyanine by sublimation (two to three times) in an argon flow at 300-400Pa; and similarly for the Cu(II), Ni(II), Pb(II), VO(II) and Zn(II) phthalocyanine complexes. [*Beilstein* **26** III/IV 4255.]

**Platinum (II) acetate (diacetatoplatinum)** [38928-89-1] **M 313.2 (monomer)**, **m 245°dec**. Purify the purple crystals by dissolving them in CHCl<sub>3</sub>, filtering, mixing with half its volume of glacial acetic acid and set aside to evaporate at room temperature. The resulting large, almost black, crystals are collected, washed with H<sub>2</sub>O and

dried in air. It dissolves in  $\text{CHCl}_3$ ,  $^*\text{C}_6\text{H}_6$  and toluene to give purple solutions. Its molecular weights (ebullioscopically) in  $\text{CHCl}_3$ ,  $^*\text{C}_6\text{H}_6$  and chlorobenzene are 950 ( $\pm 2$ ), 937 and 888, respectively, indicate that it is trimeric in solution. [Stephenson et al. *J Chem Soc* 3632 1965.]

**Platinum (II) acetylacetonate** [15170-57-7] **M 393.3, m 249-252°**. It crystallises from  $^*\text{C}_6\text{H}_6$  as yellow crystals and is dried in air or in a vacuum desiccator. [Werner *Chem Ber* **34** 2584 1901, Fernelius & Bryant *Inorg Synth* **V** 105 1957, *Beilstein* **1** H 783, **1** IV 3678.]

**Potassium 4-aminobenzoate** [138-84-1] **M 175.2**. Crystallise the aminobenzoate from EtOH. [*Beilstein* **14** II 246, **14** IV 1128.]

**Potassium antimonytartrate monohydrate** [28300-74-5] **M 333.9,  $[\alpha]_D^{20} +141^\circ$  (c 2,  $\text{H}_2\text{O}$ )**. Crystallise "tartar emetic" from water (3ml/g) between 100° and 0°. Dry it at 100°. [Chinoporos & Papathanosopolos *J Phys Chem* **65** 1643 1961, *Beilstein* **3** III 1014, **3** IV 1227.]

**Potassium benzoate** [582-25-2] **M 160.2**. Crystallise it from water (1ml/g) between 100° and 0°. [*Beilstein* **9** III 375, **9** IV 279.]

**Potassium tert-butoxide** [865-47-4] **M 112.2**. It sublimes at 220°/1mm. The last traces of *tert*-BuOH are removed by heating at 150-160°/2mm for 1 hour. It is best prepared afresh as likely impurities are *tert*-BuOH, KOH and  $\text{K}_2\text{CO}_3$  depending on its exposure to air. Its solubility at 25-26° in hexane, toluene,  $\text{Et}_2\text{O}$ , and THF is 0.27, 2.27, 4.34 and 25.0%, respectively. [Feuer et al. *J Am Chem Soc* **78** 4364, Doering & Urban *J Am Chem Soc* **78** 5938 1956, *Beilstein* **1** IV 1612.]

**Potassium citrate tribasic monohydrate (tripotassium citrate monohydrate)** [6100-05-6] **M 324.4, m 275°(dec)**. Its solubility in  $\text{H}_2\text{O}$  is 154g/100ml and it loses  $\text{H}_2\text{O}$  at 180°. [*Beilstein* **3** III 1091.]

**Potassium dihydrogen citrate** [866-83-1] **M 230.2**. It crystallises from  $\text{H}_2\text{O}$  with a solubility of 11.5% at 100°. Dry it at 80°, or in a vacuum desiccator over Sicapent. [*Beilstein* **3** I 209, **3** III 336, **3** IV 402.]

**Potassium hydrogen D-glucurate (D-tetrahydroxyadipic acid mono K salt, D-glucosaccharic acid mono K salt)** [18404-47-2] **M 248.2, m 188°(dec),  $[\alpha]_D^{20} +4.95^\circ$  (c 1.2,  $\text{H}_2\text{O}$ ),  $\text{pK}_1^{25} 5.0$  (free acid)**. Crystallise it from water. [Wolfrom & Neely *J Am Chem Soc* **75** 2778 1953, *Beilstein* **3** H 579, **3** II 378, **3** III 1119, **3** IV 1291.]

**Potassium hydrogen malate** [4675-64-3] **M 172.2**. A saturated aqueous solution at 60° is decolorised with activated charcoal, and filtered. The filtrate is cooled in a water-ice bath, and the salt is precipitated by adding EtOH. After five recrystallisations from ethanol-water mixtures, it is dried overnight at 130° in air [Eden & Bates *J Res Nat Bur Stand* **62** 161 1959]. [*Beilstein* **3** H 419, **3** III 915, **3** IV 1123.]

**Potassium hydrogen oxalate ( $\text{H}_2\text{O}$ )** [127-95-7] **M 137.1, m ~300°(dec)**. Crystallise the acid oxalate from  $\text{H}_2\text{O}$  by dissolving 20g in 100ml  $\text{H}_2\text{O}$  at 60° containing 4g of potassium oxalate, filtering and allowing to cool to 25°. The crystals, after washing three or four times with water, are allowed to dry in air. [*Beilstein* **2** III 1552.]

**Potassium hydrogen phthalate** [877-24-7] **M 204.2**. Crystallise it first from a dilute aqueous solution of  $\text{K}_2\text{CO}_3$ , then  $\text{H}_2\text{O}$  (3ml/g) between 100° and 0°. Before being used as a standard in volumetric analysis, analytical grade potassium hydrogen phthalate should be dried at 120° for 2 hours, then allowed to cool in a desiccator. [*Beilstein* **9** IV 3169.]

**Potassium hydrogen d-tartrate** [868-14-4] **M 188.2,  $[\alpha]_{\text{D}}^{20} +37.5^\circ$  (c 10, M NaOH)**. It crystallises from water (17ml/g) between 100° and 0°. Dry it at 110°. [*Beilstein* **3** IV 1222.]

**Potassium ionophore I (valinomycin)** See valinomycin in "Miscellaneous Compounds", Chapter 7.

**Potassium laurate (potassium dodecanoate)** [10124-65-9] **M 338.4**. Recrystallise potassium laurate three times from EtOH [Neto & Helene *J Phys Chem* **91** 1466 1987]. [*Beilstein* **2** H 360, **2** I 156, **2** II 318, **2** III 880, **2** IV 186.] See purification of sodium dodecanoate below.

**Potassium nonafluorobutane sulfonate** [29420-49-3] **M 338.2**. Wash it with H<sub>2</sub>O and dry it *in vacuo*. When the K salt is distilled with 100% H<sub>2</sub>SO<sub>4</sub>, it gives the *free acid* which can be distilled (**b** 105°/22mm, 210-212°/760mm) and then converted back into the pure K salt. [Gramstad & Haszeldine *J Chem Soc* 2640 1957, *Beilstein* **2** IV 818.]

**Potassium oleate** [143-18-0] **M 320.6**. Recrystallise potassium oleate from EtOH (1g/ml). [*Beilstein* **2** H 465, **2** I 196, **2** I 202, **2** II 436, **2** III 1404, **2** IV 1646.]

**Potassium oxalate** [6487-48-5] **M 184.2, m 160°(dec), d<sub>4</sub><sup>20</sup> 2.13**. Recrystallise the oxalate from hot water. It forms various K + oxalic acid salts [e.g. potassium hydrogen oxalate KH(C<sub>2</sub>O<sub>4</sub>)—see above, and KH<sub>3</sub>(C<sub>2</sub>O<sub>4</sub>)<sub>2</sub> which is, surprisingly, called potassium tetraoxalate—see below]. [*Beilstein* **2** H 513, **2** I 224, **2** II, 485, **2** III 1552, **2** IV 1823.]

**Potassium pentamethylcyclopentadienide (KCp', 1,2,3,4,5-pentamethyl-2,4-pentadiene-1-ylpotassium)** [94348-92-2] **M 174.3, m >230°**. In an efficient fume cupboard, liquid NH<sub>3</sub> (1.3L) is condensed into a flask containing freshly cut K metal (11.7g, 300mg-atoms) under an argon or N<sub>2</sub> atmosphere, stirred at -78°, and a small crystal of Fe(NO<sub>3</sub>)<sub>3</sub>·9H<sub>2</sub>O is added to the blue coloured solution. The mixture is made to warm to room temperature allowing refluxing until the blue solution turns into a light grey suspension. Cp' (40.9g, 300mmol) is added *via* a syringe, the mixture is stirred at reflux (MeOH condenser cooled at -78°) for 4 hours, and the NH<sub>3</sub> is allowed to evaporate using a stream of N<sub>2</sub>. The residue is dried *in vacuo* and stored in a N<sub>2</sub> atmosphere as it is air and moisture sensitive, and unlike other Cp'Ms it can be spontaneously flammable in air. It is more soluble in THF than Cp'Li but less soluble than Cp'Na. It is a Cp' ligand source for complexing with a large number of metals and non-metal halides. [Kohl et al. in *Organometallic Synthesis* (R. Bruce King and J.J. Eisch eds) **vol 3** 381, Elsevier Amsterdam 1986].

**Potassium phthalimide (phthalimide K salt)** [1074-82-4] **M 185.2, m >300°**. The solid may contain phthalimide and K<sub>2</sub>CO<sub>3</sub> from hydrolysis. If too much hydrolysis has occurred (this can be checked by extraction with cold Me<sub>2</sub>CO in which the salt is insoluble, evaporate the Me<sub>2</sub>CO and weigh the residue), it would be better to prepare it afresh. If little hydrolysis has occurred, then recrystallise it from a large volume of EtOH, and wash the solid with a little Me<sub>2</sub>CO and dry it in a continuous vacuum to constant weight. [Salzerg & Supriawski *Org Synth Coll Vol I* 119 1941, Raman & IR: Hase *J Mol Struct* **48** 33 1978, Dykman *Chem Ind (London)* 40 1972, IR, NMR: Assef et al. *Bull Soc Chim Fr* II 167 1979, *Beilstein* **21/10** V 270.]

**Potassium picrate** [573-83-1] **M 267.2, m 250°**. Recrystallise it from water or 95% EtOH, and dry it at room temperature in vacuum. It is soluble in 200 parts of cold water and 4 parts of boiling water. [*Beilstein* **3** III 880, **3** IV 1390.] **THE DRY SOLID EXPLODES WHEN STRUCK OR HEATED.**

**Potassium propionate** [327-62-8] **M 112.2**. Recrystallise C<sub>2</sub>H<sub>5</sub>CO<sub>2</sub>K from water (30ml/g) or 95% EtOH. [*Beilstein* **2** II 217, **2** III 516, **2** IV 701.]

**Potassium sodium tartrate tetrahydrate** [6381-59-5 (4H<sub>2</sub>O), 304-59-6 (R,R)] **M 282.3**. Recrystallise the tartrate from distilled water (1.5ml/g) by cooling to 0°. [*Beilstein* **3** IV 1223.]

**Potassium *d*-tartrate monohydrate** [921-53-9, 6381-59-5] **M 235.3, m loses H<sub>2</sub>O at 150°, d<sub>4</sub><sup>20</sup> 1.98**. Recrystallise it from distilled water (solubility: 0.4ml/g at 100°, 0.7ml/g at 14°). [*Beilstein* **3** IV 1223.]

**Potassium tetraoxalate dihydrate [oxalic acid hemipotassium salt]** [6100-20-5 (2H<sub>2</sub>O), 127-96-8] **M 254.2**. Crystallise it from water below 50°. Dry it below 60° at 760mm. [See potassium oxalate above, *Beilstein* **2** IV 1823.]

**Potassium trifluoroacetate** [2923-16-2] **M 152.1, m 140-142°**, **pK<sup>25</sup> 0.52 (for CF<sub>3</sub>CO<sub>2</sub>H)**. To purify it, dissolve the salt in trifluoroacetic acid with *ca* 2% of trifluoroacetic anhydride, filter it and evaporate it carefully to dryness (avoid over heating), and finally dry it in a vacuum at 100°. It can be recrystallised from trifluoroacetic acid (solubility in the acid is *ca* 50.1%). [Simons & Lorentzen *J Am Chem Soc* **74** 4746 1952, Hara & Cady *J Am Chem Soc* **76** 4285 1954, Watt & Muga *J Inorg Nucl Chem* **9** 166 1959, *Beilstein* **2** IV 458.]

**Pyridinium chlorochromate** [26299-14-9] **M 215.6, m 205°(dec)**. Dry it in a vacuum for 1 hour. It is not hygroscopic and can be stored for extended periods at room temperature without change. If the purity is very suspect, it can be readily prepared. [Corey & Scuggs *Tetrahedron Lett* 2647 1975, Piancatelli et al. *Synthesis* 245 1982.] (Possible **CARCINOGEN**.)

§ Available commercially on a polymer support.

**Pyridinium dichromate** [20039-37-6] **M 376.2, m 145-148°, 152-153°**. Dissolve it in the minimum volume of H<sub>2</sub>O and add 5 volumes of cold Me<sub>2</sub>CO and cool to -20°. After 3 hours the orange crystals are collected, washed with a little cold Me<sub>2</sub>CO and dried in a vacuum. It is soluble in dimethylformamide (0.9g/ml at 25°), and in H<sub>2</sub>O, and has a characteristic IR with  $\nu_{\max}$  at 930, 875, 765, 730 and 730 cm<sup>-1</sup>. [Corey & Schmidt *Tetrahedron Lett* 399 1979, Coats & Corrigan *Chem Ind (London)* 1594 1969.] (Possible **CARCINOGEN**.)

§ Available commercially on a polymer support.

**Pyronin B [di-(3,6-bis(diethylamino)xanthylium chloride) diFeCl<sub>5</sub> complex]** [2150-48-3] **M 358.9 (Fe free), m 176-178° (diFe complex), CI 45010,  $\lambda_{\max}$  555nm, pK<sup>25</sup> 7.7**. Recrystallise it from EtOH. It forms an Fe stain. [*Beilstein* **18/10** V 182.]

**Quinolinium chlorochromate** [108703-35-1] **M 265.6, m 127-130°**. It is a yellow-brown solid which is stable in air for long periods. If it has deteriorated or been kept for too long, it is best to prepare it freshly. Add freshly distilled quinoline (13ml) to a mixture of chromic acid (CrO<sub>3</sub>) (10g) and ~5M HCl (11ml of conc HCl and 10ml of H<sub>2</sub>O) at 0°. A yellow-brown solid separates, it is filtered off on a sintered glass funnel, dried for 1 hour in a vacuum, and can be stored for extended periods without serious loss in activity. It is a good oxidant for primary alcohols in CH<sub>2</sub>Cl<sub>2</sub>. [Singh et al. *Chem Ind (London)* 751 1986, method of Corey & Suggs *Tetrahedron Lett* 2647 1975, *Beilstein* **20/7** V 276.]

**Resorufin (7-hydroxy-3H-phenoxazine-3-one Na salt)** [635-78-9] **M 213.2, pK<sub>1</sub><sup>30</sup> 6.93, pK<sub>2</sub><sup>30</sup> 9.26, pK<sub>3</sub><sup>30</sup> 10.0**. Wash Resorufin with water and recrystallise it several times from EtOH. [*Beilstein* **27** II 108, **27** III/IV 2263.]

**Rhodizonic acid sodium salt (5,6-dihydroxycyclohex-5-ene-1,2,3,4-tetraone di-Na salt)** [523-21-7] **M 214.0, pK<sub>1</sub><sup>30</sup> 4.1 (4.25), pK<sub>2</sub><sup>30</sup> 4.5 (4.72)**. The *free acid* is obtained by acidifying and extracting with Et<sub>2</sub>O, drying (MgSO<sub>4</sub>), filtering, evaporating and distilling in a vacuum (**b** 155-160°/14mm). The *free acid* solidifies on cooling, and the colourless crystals can be recrystallised from tetrahydrofuran/petroleum ether or \*C<sub>6</sub>H<sub>6</sub>. It forms a *dihydrate* **m** 130-140°. The pure di Na salt is formed by dissolving the acid in 2 equivalents of NaOH and evaporating in a vacuum. It forms violet crystals which give an orange solution in H<sub>2</sub>O that is unstable for extended periods even at 0°, and should be prepared freshly before use. Salts of rhodizonic acid **cannot** be purified by recrystallisation without great loss due to conversion to crotonate, so that the original material must be prepared anew if pure salt is required. It can be washed with NaOAc solution, then EtOH, to remove excess NaOAc, dried under vacuum and stored in the dark. [UV and tautomerism: Schwarzenbach & Suter *Helv Chim Acta* **24** 617 1941, Polarography: Preisler & Berger *J Am Chem Soc* **64** 67 1942, Souchay & Taibouet *J Chim Phys* **49** C108 1952, *Beilstein* **8** H 535, **8** II 572, **8** III 4214, **8** IV 3609.]

**Ruthenium (III) acetylacetonate** [14284-93-6] **M 398.4, m 240°(dec)**. Purify the complex by recrystallisation from \*benzene. [Wilkinson *J Am Chem Soc* **74** 6146 1952, *Beilstein* **1** IV 3677.]

**Ruthenocene [bis-( $\pi$ -cyclopentadienyl)ruthenium]** [1287-13-4] **M 231.2, m 195.5°, 199-210°**. If it has to

be prepared *ab initio*, then in a flask (500ml), under continuous flushing with N<sub>2</sub>, and stirring, dry 1,2-dimethoxyethane (300ml) is added followed by Na (7.2g, 0.312 g.atom) cut in small pieces then cyclopentadiene (31.0ml, 0.376 mole; see [542-92-7], prepared from the dimer, [cf. Wilkinson *Org Synth Coll Vol IV* 475 1963] is added dropwise. After H<sub>2</sub> evolution stops, the mixture is kept just below boiling for 1-2 hours until all the Na has dissolved. If some metal remains undissolved, cool, and add a few milliliters of cyclopentadiene, and heat again until complete dissolution of the metal occurs. To this solution of sodium pentadienide [see also 4989-82-1], under N<sub>2</sub> and stirring, is added a mixture of RuCl<sub>3</sub> (14.6g, 0.07 mole) and Ru metal (2.4g, 0.024 g.atom), and heated just below reflux for 80 hours. The solvent is removed under reduced pressure, the flask is filled with N<sub>2</sub>, the solid residue is transferred (use a dry-box) to a sublimator, and sublimed at 0.1mm and 130° (bath temperature). It is advantageous to use a Dry-ice sublimation finger. The sublimate is dissolved in \*C<sub>6</sub>H<sub>6</sub>, passed through a column of activated Al<sub>2</sub>O<sub>3</sub> (2.5 x 30cm), and evaporated to give pure *non-pyrophoric* ruthenocene (12.2-15.1g, 56-69%), m 199-200°. The residue from sublimation is, however, *pyrophoric*, but this residue can be destroyed by adding H<sub>2</sub>O under N<sub>2</sub>. It can be further sublimed in high vacuum at 120°. It forms yellow crystals which can be recrystallised from CCl<sub>4</sub> as transparent plates. [Bublitz et al. *Org Synth Coll Vol V* 1001 1973, Wilkinson *J Am Chem Soc* 74 6146 1952, *Beilstein* 16 IV 1833.] It is generally less reactive than ferrocene towards electrophiles.

Two other alternative syntheses have been adopted using Schlenk techniques in an atmosphere of N<sub>2</sub> or argon, with dry and de-oxygenated solvents. In the *first*, [Sn-*n*-Bu<sub>3</sub>(C<sub>5</sub>H<sub>5</sub>)] (32.0g, 90mmol, prepared from Sn-*n*-Bu<sub>3</sub>Cl, 1461-22-9, and LiC<sub>5</sub>H<sub>5</sub> as per Davison & Rakita *Inorg Chem* 9 289 1970) is added to a suspension of [RuCl<sub>2</sub>(COD)<sub>n</sub>] (8.4g, 30mmol, see [50982-12-2]) in EtOH (150ml), and the mixture is stirred and refluxed for 48 hours whereby the Ru polymer dissolved. The hot solution is filtered in air, and on cooling the pale yellow *ruthenocene* that crystallises out is filtered off, washed thoroughly with cold EtOH (2 x 20ml) then Et<sub>2</sub>O (2 x 20ml), and recrystallised to analytical purity (5.2g, 75%) from hot EtOH or Me<sub>2</sub>CO. Its <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) has δ at 4.55 (s, C<sub>5</sub>H<sub>5</sub>), and its <sup>13</sup>C NMR (125MHz, CDCl<sub>3</sub>) has δ at 20.1 (s, C<sub>5</sub>H<sub>5</sub>) (any *n*-butyl signals from impurities would be at their usual shifts). In the *second*, [RuCl<sub>2</sub>(COD)<sub>n</sub>] (1.4g, 50mmol) and cyclopentadienyl thallium (2.7g, 10mmol, see [34822-90-7]) in dimethoxyethane (~50ml) are refluxed for 1 hour, filtered hot (in air), the solvent is removed *in vacuo*, the residue is extracted with Et<sub>2</sub>O (4 x 20ml), filtered, evaporated, and the residue is recrystallised and/or sublimed, as above, to give pure *ruthenocene* (0.9g, 78%). [Albers et al. *Organometallics* 5 2321 1986.] The Raman spectrum of the melt has ν<sub>max</sub> at 160w, 325s(p), 396m, 590vw, 815m(p), 830w, 1060m, 1098s(p), 1195vw 1360w, 1410m cm<sup>-1</sup> (p is with polarized light) [Lokshin et al. *J Organomet Chem* 124 293 1977].

**Silver acetate** [563-63-3] M 166.9, pK<sup>25</sup> >11.1 (for aquo Ag<sup>+</sup> hydrolysis). Shake AgOAc with acetic acid for three days, and the process is repeated with fresh acid. The solid is then dried in a vacuum oven at 40° for 48 hours. It has also been recrystallised from water containing a trace of acetic acid, and dried in air. Store it in the dark. [*Beilstein* 2 IV 112.]

**Silver lactate** [128-00-7] M 196.9, m ~ 100°. Recrystallise it from H<sub>2</sub>O by adding EtOH. The solid is collected, washed with EtOH, then Et<sub>2</sub>O, and dried at 80° to give the *dihydrate*. It is a white powder soluble in 15 parts of H<sub>2</sub>O but only slightly soluble in EtOH. Store it in the dark. [Engelhardt & Maddrell *Justus Liebigs Ann Chem* 63 89 1847, Karrer et al. *Helv Chim Acta* 2 251 1919, *Beilstein* 3 III 464.]

**Silver tetrafluoroborate** [AgBF<sub>4</sub>] [14104-20-2] M 194.7, m 70-73°, dec >200° (to form a silver mirror). The anhydrous salt has been freshly prepared under dry N<sub>2</sub>, by suspending AgF (34g, 268mmol) in dry MeNO<sub>2</sub> (30.5ml) while BF<sub>3</sub> is passed through with stirring. The temperature rises to 60°, and is maintained by occasional dipping into an ice bath. After about 30 minutes most of the AgF dissolves and ~280mmoles of BF<sub>3</sub> will have dissolved. Excess BF<sub>3</sub> is removed by bubbling dry N<sub>2</sub> for 1 hour through the murky solution which is then filtered (through a sintered glass frit, under N<sub>2</sub>), and the MeNO<sub>2</sub> is distilled off at 70°/5mm into a liquid N<sub>2</sub> trap. The off-white solid is ground in a mortar under dry *n*-pentane in a glove box flushed with dry N<sub>2</sub>, and the solvent is evaporated under N<sub>2</sub> at ~25°. The residual salt (237mmole, 88.5%) is pumped dry at 70° and stored under *n*-pentane. [Olah & Quin *J Inorg Nucl Chem* 14 295 1960.] The synthesis has been carried out in \*C<sub>6</sub>H<sub>6</sub>

with which it forms a 1:2 complex [Heyns & Paulsen *Angew Chem* **72** 349 1960], although 1:1 [Meerwein et al. *Arch Pharm* **291** 541 1958] and 2:3 complexes [Sharp & Sharpe *J Chem Soc* 1855 1956] were reported.  $\text{AgBF}_4$  is very soluble in  $\text{H}_2\text{O}$ ,  $\text{Et}_2\text{O}$ , toluene,  $\text{MeNO}_3$ , moderately soluble in  $^*\text{C}_6\text{H}_6$ , cyclohexene but insoluble in pentane and cyclohexane. The solubility in unsaturated hydrocarbons is attributed to  $\pi$ -bond complexing. A solution obtained by dissolving  $\text{Ag}_2\text{O}$  (1.0g) in 45% fluoroboric acid (7.2g) has been used for the synthesis of a tropone [Birch & Keeton *J Chem Soc (C)* 09 1968].

The salt is used for making tetrafluoroborate salts of complexes and other salts (see examples in this chapter) by displacing chloride from precious metal complexes, e.g.  $\text{Ru}(\text{Me}_2\text{SO}_4)\text{Cl}_2$  with excess  $\text{AgBF}_4$  under  $\text{N}_2$  in refluxing  $\text{EtOH}$  (ca 1 hour) and repeatedly recrystallising from  $\text{H}_2\text{O}/\text{EtOH}$  to give a 60% yield of  $[\text{Ru}(\text{Me}_2\text{SO}_4)(\text{H}_2\text{O})_2]\text{BF}_4$ , which is used for the oxidation of alcohols by persulfate [Bressan et al. *J Mol Catal* **79** 85 1993]. It provides a silver-ion template for improved macrolactamisation of linear dipeptides, e.g. to form cyclohexadepsipeptides [Yanjie et al. *Synlett* 1901 2007].

**Silver trifluoroacetate** [2966-50-9] **M 220.9, m 251-255°**. Extract the salt (Soxhlet) with  $\text{Et}_2\text{O}$ . The extract is filtered and evaporated to dryness, then the powdered residue is completely dried in a vacuum desiccator over silica gel. Its solubility in  $\text{Et}_2\text{O}$  is 33.5g in 750ml. It can be recrystallised from  $^*\text{C}_6\text{H}_6$  (solubility is: 1.9g in 30ml of  $^*\text{C}_6\text{H}_6$ , and 33.5g will dissolve in 750ml of anhydrous  $\text{Et}_2\text{O}$ ). [Traynham & Dehn *J Org Chem* **23** 1545 1958, Haszeldine *J Chem Soc* 584 1951.] Store it in the dark. It is also soluble in trifluoroacetic acid (15.2% at 30°), toluene, *o*-xylene and dioxane [Hara & Cady *J Am Chem Soc* **76** 4285 1954]. [Beilstein **2** IV 461.]

**Sodium acetate** (anhydrous) [127-09-3] **M 82.0, m 324°, d<sub>4</sub><sup>20</sup> 1.53**. Crystallise it from acetic acid and keep it under vacuum for 10 hours at 120°. Alternatively, it crystallises from aqueous  $\text{EtOH}$  as the *trihydrate*. This material can be converted to *anhydrous* salt by heating slowly in a porcelain, nickel or iron dish, so that the salt liquefies. Steam is evolved and the mass again solidifies. Heating is now increased so that the salt melts again. (NB: if it is heated too strongly, the salt can char; avoid this.) After several minutes, the salt is allowed to solidify and is cooled to a convenient temperature (in a desiccator) before being powdered and bottled. The water content should now be less than 0.02%. [Beilstein **2** II 113, **2** III 184, **2** IV 109.]

**Sodium acetylacetonate** [ $\text{Na}(\text{acac})$ , 2,4-pentanedione ion (1-) sodium (1:1)] [15435-71-9] **M 122.1, m 217-219° and is quite stable to heating below 191°, d<sub>4</sub><sup>20</sup> 1.213**. It is prepared by two different procedures. In the *first*, 2,4-pentanedione (100g, 1.0 mole) in  $\text{H}_2\text{O}$  or  $\text{EtOH}$  is slowly added to  $\text{NaOH}$  (40g, 1.0 mole) in the minimum amount of  $\text{H}_2\text{O}$  while keeping the temperature below 70°. On cooling  $\text{Na}(\text{acac})$  (98g, 80%) crystallises out. Note that addition of solid  $\text{NaOH}$  resulted in a product containing  $\text{NaOH}$ . Recrystallisation from aqueous  $\text{EtOH}$ ,  $\text{Me}_2\text{CO}$  or  $\text{EtOAc}$  gives pearly plates of the *dihydrate* (most probably a coordinated compound) which is soluble  $\text{EtOAc}$  but not in toluene. It gives off  $\text{H}_2\text{O}$  on heating, and the anhydrous form (salt form ?) is no longer soluble in  $\text{EtOAc}$ .

In the *second* procedure, 2,4-pentanedione (50g, 0.5 mole) is added to  $\text{Na}$  metal (11.5g, 0.5 mole) in toluene (100ml) while the temperature is maintained at 100° to keep the metal in the molten state (CARE, as the experiment may be dangerous;  $\text{N}_2$  should be flushing through the apparatus until the metal has dissolved).  $\text{Na}(\text{acac})$  (101g, 83%) separates during the reaction. Note that using lower temperatures provide  $\text{Na}(\text{acac})$  contaminated with  $\text{Na}$ . The product can be recrystallised from  $\text{EtOH}$  (see above) or by dissolving in  $\text{EtOH}$  and adding  $^*\text{C}_6\text{H}_6$ . Its solubility in *n*-hexane, cyclohexane, or  $^*\text{C}_6\text{H}_6$  is less than 0.01% at 70°. However, although quite insoluble in  $^*\text{C}_6\text{H}_6$  or toluene, a suspension of  $\text{Na}(\text{acac})$  will dissolve in these solvents on warming if excess of 2,4-pentanedione, or other related dicarbonyl compounds, e.g. benzoylacetone or ethyl acetoacetate, are added. Note that on heating the anhydrous or hydrated material decomposition and charring occurs if the sample in a soft glass tube is used. However, if the tube is placed in a bath at ~212° and the temperature is raised slowly, a sharp and reproducible melting point (217-219°) can be obtained with very little decomposition. [Hatch & Sutherland *J Org Chem* **13** 249 1948, Sidgwick & Brewer *J Chem Soc* **127** 2379 1925.] The ratios of mole of decomposition product/mole of "salt" after 50 hours when heated under  $\text{N}_2$  at 27° and 191° are both 0.07, i.e. it is reasonably stable at high temperatures. [Charles & Pawlikowski *J Phys Chem* **62** 440 1958.] Its UV ( $\text{EtOH}$ ) has  $\lambda_{\text{max}}$  at 277 ( $\epsilon$  800 x 10<sup>3</sup>) nm [Hatch & Sutherland *J Org Chem* **13** 249 1948]; and its electrical molecular conductivities (dilution) are: 5.794 (8), 8.254 (16), 11.187 (32), 14.426 (64), 18.116 (128), 21.82 (256), 25.636 (512), 29.927 (1024), and do not approach complete dissociation when nearing maximum dilution as hydrolysis may well be occurring [White *J Chem Soc* 1413 1928]. Like  $\text{K}(\text{acac})$  and  $\text{Cs}(\text{acac})$ ,  $\text{Na}(\text{acac})$  is

unstable in aqueous solution or in moist conditions and is decomposed by hot H<sub>2</sub>O into Me<sub>2</sub>CO and the alkali acetate [Morgan & Moss *J Chem Soc* **105** 189 1914]. Store it in a dry atmosphere. [Marchi *Inorg Synth* **II** 10 1946, Frenelius & Bryant *Inorg Synth* **V** 105 1957.]

**Sodium acetylide** [1066-26-8] **M 48.0**. It disproportionates at *ca* 180° to sodium carbide. It sometimes contains diluents, e.g. xylene, butyl ether or dioxane that can be removed by filtration followed by a vacuum at 65-60°/5mm. *Alternatively*, the acetylide is purged with HC≡CH at 100-125° to remove diluent. NaC<sub>2</sub>H adsorbs 2.2x, 2.0x and 1.6x its wt of xylene, butyl ether and dioxane, respectively. Powdered NaC<sub>2</sub>H is yellow or yellow-gray in colour and is relatively stable. It can be heated to *ca* 300° in the absence of air. Although no explosion or evolution of gas occurs, it turns brown due to disproportionation. At 170-190° in air it ignites slowly and burns smoothly. At 215-235° in air it “flash-ignites” and burns quickly. It can be dropped into a *slight* excess of H<sub>2</sub>O without flashing or burning, but vigorous evolution of HC≡CH (**HIGHLY FLAMMABLE IN AIR**) occurs. A sample had been stored in the absence of air for one year without deterioration. Due to the high flammability of HC≡CH, the salt should be stored dry and should be treated with care. After long storage, NaC≡CH can be redissolved in liquid NH<sub>3</sub> and used for the same purposes as the fresh material. However, it may be slightly turbid due to the presence of moisture. [Rutledge *J Org Chem* **22** 649 1957, Greenlee & Henne *J Am Chem Soc* **77** 5013 1955, Campbell & Campbell *Inorg Synth* **II** 76, 81 1946, *Org Synth* **30** 15 1950, *Beilstein* **1 H** 238.] It is available commercially under N<sub>2</sub> in Sure/Seal bottles as an 18 wt% solution in xylene/mineral oil. See “Aliphatic Compounds”, Chapter 4, for its preparation.

**Sodium alginate (Algin)** [9005-38-3]. Free Algin from heavy metal impurities by treatment with ion-exchange resins (Na<sup>+</sup>-form), or with a dilute solution of the sodium salt of EDTA. *Alternatively*, dissolve it in 0.1M NaCl, centrifuge and fractionally precipitate it by gradual addition of EtOH or 4M NaCl. The resulting gels are centrifuged off, washed with aqueous EtOH or acetone, and dried under vacuum. [Büchner et al. *J Chem Soc* 3974 1961.]

**Sodium 4-aminobenzoate** [555-06-6] **M 159.1**. Recrystallise it from water. [Hermann *Helv Chim Acta* **9** 786 1926, *Beilstein* **14** II 247.]

**Sodium 4-aminosalicylate dihydrate** [6018-19-5] **M 175.1**. Recrystallise it from water at room temperature (2ml/g) by adding acetone and cooling. [*Beilstein* **14** III 1436, **14** IV 1969.]

**Sodium antimonyl tartrate** [34521-09-0] **M 308.8**. It crystallises from water. [*Beilstein* **3** III 1014, **3** IV 1227.]

**Sodium barbitone (sodium 5,5-diethylbarbiturate)** [144-02-5] **M 150.1**, pK<sub>1</sub><sup>25</sup> 3.99, pK<sub>2</sub><sup>25</sup> 12.5 (**barbituric acid**). Crystallise it from water (3ml/g) by adding an equal volume of EtOH and cooling to 5°. Dry it under vacuum over P<sub>2</sub>O<sub>5</sub>. [*Beilstein* **24** III/IV 1904.]

**Sodium benzoate** [532-32-1] **M 144.1**. Crystallise it from EtOH (12ml/g). [*Beilstein* **9** IV 27.]

**Sodium tert-butoxide** [865-48-5] **M 96.1**. It sublimes at 180°/1mm. Its solubility in *tert*-BuOH is 0.208M at 30.2° and 0.382M at 60°, and it is quite soluble in tetrahydrofuran (32g/100g). It should not be used if it has a brown colour. [Feuer *J Am Chem Soc* **78** 4364 1956, Hurd *Inorg Synth* **I** 87 1939, IR: Seubold *J Org Chem* **21** 156 1956, *Beilstein* **1** IV 1609.]

**Sodium butyrate** [156-54-7] **M 110.1**. Prepare it by neutralising the acid with Na<sub>2</sub>CO<sub>3</sub>, and recrystallising it from EtOH. [*Beilstein* **2** IV 779.]

**Sodium carboxymethylcellulose** [9004-32-4]. Dialyse it for 48 hours against distilled water and freeze-dry if a solid is required.

**Sodium decanoate (sodium caproate)** [1002-62-6] **M 194.2**. Neutralise sodium hydroxide by adding a slight excess of free decanoic acid and recovering the excess acid by Et<sub>2</sub>O extraction. The salt is recrystallised from

the aqueous solution by adding pure acetone and repeating the steps several times, then drying the salt in an oven at *ca* 110° [Chaudhury & Awuwallia *Trans Faraday Soc* 77 3119 1981]. [Beilstein 2 IV 1041.]

**Sodium deoxycholate monohydrate** [302-95-4, 145224-92-4] **M 432.6**,  $[\alpha]_D^{20} +48^\circ$  (c 1, EtOH). Recrystallise it from EtOH and dry it in an oven at 100°. The solution is freed from soluble components by repeated extraction with acid-washed charcoal. [Beilstein 10 IV 1608.]

**Sodium diethyloxaloacetate** [63277-17-8, 40876-98-0] **M 210.2**. Extract it several times with boiling Et<sub>2</sub>O (until the solvent remains colourless), and then the residue is dried in air. [Beilstein 3 H 789.]

**Sodium diformylamide** [18197-26-7] **M 95.0**. Grind the amide under dry tetrahydrofuran (fumehood), filter and wash it with this solvent, then dry it *in vacuo*. It is soluble in EtOH and H<sub>2</sub>O but insoluble in Et<sub>2</sub>O and petroleum ether. [IR and preparation: Rakshit *J Chem Soc* 103 1557 1913, Yinglin & Hongwen *Synthesis* 122 1990, Allenstein & Beyl *Chem Ber* 100 355 1967, Allenstein et al. *Chem Ber* 102 4089 1969, Beilstein 2 II 22.]

**Sodium p-dimethylaminoazobenzene-o'-carboxylate (Methyl Red)** [845-10-3] **M 291.2**. It can be precipitated from aqueous solution as the free acid which is recrystallised from 95% EtOH, then reconverted to the sodium salt. In H<sub>2</sub>O it is pink at pH 4.2 and yellow at pH 6.2 and has UV with  $\lambda_{\max}$  at 437nm. *Methyl Red Hydrochloride* [63451-28-5] has **M 305.8**, **m 175°(dec)** and UV with  $\lambda_{\max}$  at 493nm. [Beilstein 16 H 329.]

**Sodium p-dimethylaminoazobenzene-p'-carboxylate** [845-46-5] **M 219.2**. It has been precipitated from aqueous solution as the free acid which was recrystallised from 95% EtOH, then reconverted to the sodium salt. [Beilstein 16 H 329.]

**Sodium dodecanoate (sodium laurate)** [629-25-4] **M 222.3**, **pK<sup>20</sup> 5.3 (COOH)**. Neutralise it by adding a slight excess of dodecanoic acid and removing it by ether extraction. The salt is recrystallised from the aqueous solution by adding pure Me<sub>2</sub>CO and repeating the process (see sodium decanoate above). It has also been recrystallised from MeOH. [Beilstein 2 IV 1085.]

**Sodium ethoxide** [141-52-6] **M 68.1**. It is a *hygroscopic* powder which should be stored under N<sub>2</sub> in a cool place. A likely impurity is EtOH which can be removed by warming at 60-80° under high vacuum. It is hydrolysed by H<sub>2</sub>O to yield NaOH and EtOH. Other impurities, if kept in air for long periods are NaOH and Na<sub>2</sub>CO<sub>3</sub>. In this case the powder cannot be used if these impurities affect the reactivity, and a fresh sample would be acquired. It is prepared by adding Na to absolute EtOH under N<sub>2</sub> and evaporated *in vacuo*, or used *in situ*. [IR: Seubold *J Org Chem* 21 156 1956]. [Beilstein 1 H 311, 1 IV 1289.]

**Sodium ethylmercurithiosalicylate (Thimerosal)** [54-64-8] **sM 404.8**, **m ~230°**. Recrystallise this antibacterial from EtOH/Et<sub>2</sub>O. **HIGHLY TOXIC**. [Trikojus *Nature* 158 472 1940, Beilstein 10 III 213.]

**Sodium fluoroacetate (FCH<sub>2</sub>CO<sub>2</sub>Na)** [62-74-8] **M 100.0**, **m 200-205°(dec)**. It is a free-flowing white **HIGHLY POISONOUS** powder which is purified by dissolving it in *ca* 4 parts of H<sub>2</sub>O, and the pH is checked. If it is alkaline, add a few drops of FCH<sub>2</sub>CO<sub>2</sub>H to make the solution just acidic. Evaporate (fumehood) on a steam bath until crystals start to separate, cool and filter the solid off. More solid can be obtained by adding EtOH to the filtrate. Dry it at 100° *in vacuo*. The *p-nitrobenzyl ester* crystallises from EtOH with **m 76°**. The free acid interferes with the citric acid cycle. [Saunders & Stacey *J Chem Soc* 1778 1948, Beilstein 2 IV 446.]

**Sodium formate (anhydrous)** [141-53-7] **M 68.0**, **m 253°**, **d<sub>4</sub><sup>20</sup> 1.92**. A saturated aqueous solution at 90° (0.8ml water/g) is filtered and allowed to cool slowly. (The final temperature should be above 30° to prevent formation of the hydrate.) After two such crystallisations, the crystals are dried in an oven at 130°, then under high vacuum. [Westrum et al. *J Phys Chem* 64 1553 1960, Roecker & Meyer *J Am Chem Soc* 108 4066 1986.] The salt has also been recrystallised twice from 1mM DTPA (diethylenetriaminepentaacetic acid, which was recrystallised 4x from MilliQ water and dried in a vacuum), then twice from water [Bielski & Thomas *J Am Chem Soc* 109 7761 1987]. [Beilstein 2 IV 3.]

**Sodium D-gluconate** [527-07-1] M 218.1, m 200-205°dec,  $[\alpha]_{546}^{20} +14^{\circ}$ ,  $[\alpha]_{D}^{20} +12^{\circ}$  (c 20, H<sub>2</sub>O). Crystallise it from a small volume of H<sub>2</sub>O (solubility is 59g/100ml at 25°), or dissolve it in H<sub>2</sub>O and add EtOH since it is sparingly soluble in EtOH. It is insoluble in Et<sub>2</sub>O. It forms a Cu complex in alkaline solution and a complex with Fe in neutral solution. [Sawyer & Bagger *J Am Chem Soc* **81** 5302 1959, *Beilstein* **3** I 188.]

**Sodium glycochenodeoxycholate** [16564-43-5] M 472.6, m 245-250°,  $[\alpha]_{D}^{23} +41.2^{\circ}$  (c 1, H<sub>2</sub>O). Dissolve it in EtOH, filter it and concentrate it to crystallisation, and recrystallise from a little EtOH. It also recrystallises from EtOH/Et<sub>2</sub>O. [*Beilstein* **10** IV 1611.]

**Sodium glycocholate** [863-57-0] M 488.6, m 230-240°,  $[\alpha]_{D}^{23} +29.2^{\circ}$  (c 1, H<sub>2</sub>O). Dissolve it in EtOH, filter it and concentrated to crystallisation, and recrystallise from a little EtOH. It also crystallises from 95% EtOH/Et<sub>2</sub>O. [*Beilstein* **3** IV 573.]

**Sodium glycolate dehydrate (sodium hydroxyacetate)** [2836-32-0] M 98.0, pK<sup>25</sup> 3.83 (for acid). Precipitate it from aqueous solution by adding EtOH and dry it in air. Also recrystallise it from H<sub>2</sub>O, where its solubility is 38% at 0° and 61% at 100°. [*Beilstein* **3** III 370, **3** IV 573.]

**Sodium hydrogen diglycollate (2,2'-oxydiacetic acid monosodium salt)** [50795-24-9] M 156.1, pK<sub>1</sub><sup>20</sup> 2.70, pK<sub>2</sub><sup>20</sup> 4.22 (for acid). Crystallise it from hot water (7.5ml/g) by cooling to 0° with constant stirring. The crystals are filtered off on to a sintered-glass funnel and dried at 110° overnight. Its solubility in water is 2.6% (at 0°) and 20% (at 90°). [*Beilstein* **3** III 377, **3** IV 577.]

**Sodium hydrogen oxalate dihydrate** [1186-49-8] M 130.0, m 100°(loses H<sub>2</sub>O), b 200°(dec). Crystallise it from hot water (5ml/g) by cooling. Its solubility in H<sub>2</sub>O is 1.7% at 15.5° and 21.3% at boiling. [Souhay & Lensen *Justus Liebigs Ann Chem* **99** 34 1856, *Beilstein* **2** H 513, **2** I 223, IV 1817.]

**Sodium hydrogen succinate** [2922-54-5] M 140.0. Crystallise it from hot water and dry it at 110°. Its solubility in H<sub>2</sub>O is 17% at 0°, 40% at 25° and 86% at 75°. [Marshall & Bain *J Chem Soc* **97** [I] 1074, 1084 1910, Stokes **70** 1945 1948, *Beilstein* **2** H 606, **2** I 262, **2** III 16557, **2** IV 1911.]

**Sodium hydrogen d-tartrate** [526-94-3] M 190.1, m 100°(loses H<sub>2</sub>O), m 253°dec,  $[\alpha]_{546} +26^{\circ}$  (c 1, H<sub>2</sub>O) and  $[\alpha]_{D}^{20} +24^{\circ}$  (c 1, H<sub>2</sub>O). It crystallises from warm water (10ml/g) by cooling to 0°. [*Beilstein* **3** IV 1219.]

**Sodium 2-hydroxy-4-methoxybenzophenone-5-sulfonate** [6628-37-1] M 330.3. Crystallise it from MeOH and dry it under vacuum.

**Sodium p-hydroxyphenylazobenzene-p'-sulfonate** [2623-36-1] M 288.2. Recrystallise it from 95% EtOH.

**Sodium ionophore I (ETH 227) (N,N',N''-triheptyl-N,N',N''-trimethyl-4,4',4''-propylidyne-tris(3-oxabutamide)** [61183-76-4] M 642.0. It is purified (ca 200mg) by TLC on Kieselgel F<sub>254</sub> with CHCl<sub>3</sub>/Me<sub>2</sub>CO (1:1) as solvent, followed by HPLC (50mg) with an octadecyltrimethylsilane modified column (Mercksorb SI 100, 10µm) [IR, NMR, MS: Güggi et al. *Helv Chim Acta* **59** 2417 1976]. [See Simon & Carafoli *Methods Enzymol* **56** 439 1977.]

**Sodium ionophore V (ETH 4120) [4-octadecanoyloxymethyl-N,N,N',N''-tetraacyclohexyl-1,2-phenylenedioxydiacetamide]** [129880-73-5] M 849.3. Purify it by recrystallisation from EtOAc [Preparation and properties: Géhrig et al. *Anal Chim Acta* **233** 295 1990].

**Sodium ionophore VI {bis[(12-crown-4)methyl]dodecyl methyl malonate}** [80403-59-4] M 662.9. Purify it by gel permeation or column chromatography. [Preparation and NMR data: Shono et al. *J Electroanal Chem* **132** 99 1982, Cram & Cram *Science* **183** 803 1974.]

**Sodium RS-mandelate** [114-21-6] M 174.1, pK<sup>25</sup> 3.41 (for the acid). It crystallises from 95% EtOH and is dried in a vacuum. [Ross & Morrison *J Chem Soc* 1019 1933, Banks & Davies *J Chem Soc* 74 1938, *Beilstein* **10** III 450.]

**Sodium methoxide** [124-41-4] **M 54.0**. It behaves in the same way as sodium ethoxide. It is *hygroscopic* and is hydrolysed by moist air to NaOH and MeOH. Material that has been kept under N<sub>2</sub> should be used. If erratic results are obtained, even with recently purchased NaOMe, it should be freshly prepared thus: Clean Na (37g) cut in 1-3g pieces is added in small portions to stirred MeOH (800ml) in a 2L three-necked flask equipped with a stirrer and a condenser with a drying tube. After all the Na has dissolved, the MeOH is removed by distillation under vacuum, and the residual NaOMe is dried by heating at 150° under vacuum and kept under dry N<sub>2</sub> until required [Burness *Org Synth* **39** 51 1959]. [Beilstein **1** IV 1227.]

**Sodium monensin** [22373-78-0] **M 693.8**. Crystallise it from EtOH/H<sub>2</sub>O [Cox et al. *J Am Chem Soc* **107** 4297 1985].

**Sodium oleate** [143-19-1] **M 304.4, m 233-235°**. Sodium oleate crystallises from EtOH, and is dried in an oven at 100°. [Beilstein **2** H 465, **2** I 201, **2** II 434, **2** III 1405, **2** IV 1645.]

**Sodium oxalate** [62-76-0] **M 134.0, m 250-270°(dec), d<sub>4</sub><sup>20</sup> 2.34**. It crystallises from hot water (16ml/g) by cooling to 0°. Before use as a volumetric standard, analytical grade quality sodium oxalate should be dried for 2 hours at 120° and allowed to cool in a desiccator. [Beilstein **2** IV 1819.]

**Sodium palmitate** [408-35-5] **M 278.4, m , 270°, 285-201°**. It crystallises from EtOH and is dried in an oven. [Beilstein **2** IV 1157.]

**Sodium pentamethylcyclopentadienide (NaCp', 1,2,3,4,5-pentamethyl-2,4-pentadiene-1-ylsodium)** [40585-51-1] **M 158.2, m >230°**. It is prepared, handled and stored as for KCp' but is not as flammable. A 0.5M solution in THF is commercially available. It is a Cp' ligand source for complexing with a large number of metals and non-metal halides. [Kohl et al. in *Organometallic Synthesis* (R. Bruce King and J.J. Eisch eds) **Vol 3** 381, Elsevier Amsterdam 1986].

**Sodium phenoxide** [139-02-6, 156150-40-2 (3H<sub>2</sub>O)] **M 116.1, m 61-64°**. The ground powder is washed with Et<sub>2</sub>O, then heated at 60°/1mm for 12 to 24 hours to remove any free phenol and solvent. [Kornblum & Lurie *J Am Chem Soc* **81** 2710 1959, Beilstein **6** I 718.]

**Sodium phenylacetate** [114-70-5] **M 158.1**. Its aqueous solution is evaporated to crystallisation on a steam bath; the crystals are washed with absolute EtOH and dried under vacuum at 80°. [Beilstein **9** IV 1614.]

**Sodium o-phenylphenolate tetrahydrate** [132-27-4] **M 264.3**. Crystallise the salt from acetone and dry it under vacuum at room temperature. [Beilstein **16** IV 4600.]

**Sodium phenylpyruvate** [114-76-1] **M 186.1, m >300°**. The salt should have no OH broad bands in the IR at ~3000cm<sup>-1</sup>. If these are present, then they are due either to water contamination or to the presence of free acid. For the first case dry the solid thoroughly in a vacuum over P<sub>2</sub>O<sub>5</sub>, and in the latter case wash the salt well with Et<sub>2</sub>O *in vacuo* till free of acid. *Alternatively*, add a slight excess of the free acid (cf p 375, see[156-06-9]) in EtOH to ethanolic NaOH, evaporate to dryness and extract excess acid from the salt with dry Et<sub>2</sub>O. [Beilstein **10** I 325.]

**Sodium phytate monohydrate [myo-inositolhexakis(H<sub>2</sub>PO<sub>4</sub>) Na salt]** [14306-25-3] **M 857.9**. Crystallise sodium phytate from hot water. [Beilstein **6** IV 7927.]

**Sodium polyacrylate (NaPAA)** [9003-04-7]. Commercial polyacrylamide is first neutralised with an aqueous solution of NaOH, and the polymer is precipitated with acetone. The precipitate is redissolved in a small amount of water and freeze-dried. The polymer is then repeatedly washed with EtOH and water to remove traces of low-molecular-weight material, and finally dried in vacuum at 60° [Vink *J Chem Soc, Faraday Trans 1* **75** 1207 1979]. It has also been dialysed overnight against distilled water, then freeze-dried.

**Sodium poly ( $\alpha$ -L-glutamate).** Wash it with acetone, dry it in a vacuum, dissolve it in water and precipitate it with isopropanol at 5°. Impurities and low-molecular-weight fractions are removed by dialysis of the aqueous solution for 50 hours, followed by ultrafiltration through a filter impermeable to polymers of molecular weights greater than  $10^4$ . The polymer is recovered by freeze-drying. [Mori et al. *J Chem Soc, Faraday Trans 1* 2583 1978.]

**Sodium propionate** [137-40-6] **M 96.1, m 287-289°.** Recrystallise it from H<sub>2</sub>O (solubility 10%) and dry by heating at 100° for 4 hours. The solubility of the anhydrous salt in MeOH is 13% at 15° and 13.77% at 68°. It is insoluble in \*C<sub>6</sub>H<sub>6</sub> and Me<sub>2</sub>CO. [Henstock *J Chem Soc* 1341 1934, *Beilstein 2* IV 701.]

**Sodium stearate** [822-16-2] **M 306.6.** It is better to prepare it by adding a slight excess of octadecanoic acid to ethanolic NaOH, evaporating and extracting the residue with dry Et<sub>2</sub>O. [*Beilstein 2* III 1003.]

**Sodium R-tartrate dihydrate** [6106-24-7] **M 230.1, m 120°(loses H<sub>2</sub>O), d<sub>4</sub><sup>20</sup> 1.82, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +26° (c 1, H<sub>2</sub>O).** It crystallises from warm dilute aqueous NaOH on cooling. [*Beilstein 3* H 524.]

**Sodium trifluoroacetate** [2923-18-4] **M 136.0, m 206-210°(dec), pK<sup>25</sup> 0.52 (for CF<sub>3</sub>CO<sub>2</sub><sup>-</sup>).** A possible contaminant is NaCl. The solid is treated with CF<sub>3</sub>CO<sub>2</sub>H and evaporated twice. Its solubility in CF<sub>3</sub>CO<sub>2</sub>H is 13.1% at 29.8°. The residue is crystallised from dilute EtOH, and the solid is dried in vacuum at 100°. [Hara & Cady *J Am Chem Soc* 76 4285 1954.] It can be precipitated from EtOH by adding dioxane, then recrystallising several times from hot absolute EtOH. Dry it at 120-130°/1mm. [*Beilstein 2* IV 461.]

**Stannous bis-cyclopentadienyl** [26078-96-6] **M 248.9.** Purify it by vacuum sublimation. Handle and store it under dry N<sub>2</sub>. The related thallium and indium compounds are similarly purified.

**Strontium acetate** [543-94-2] **M 205.7, d 2.1, pK<sup>25</sup> 13.0 (for aquo Sr<sup>2+</sup> hydrolysis).** Crystallise it from AcOH, then dry it under vacuum for 24 hours at 100°. [*Beilstein 2* II 91.]

**Strontium lactate trihydrate** [29870-99-3] **M 319.8, m 120°(loses 3H<sub>2</sub>O).** It crystallises from aqueous EtOH. [*Beilstein 3* IV 633.]

**Strontium oxalate monohydrate** [814-95-9] **M 193.6, m 150°.** It crystallises from hot water on cooling. The solubility at ~25° in H<sub>2</sub>O is 1g/20L, and in 3.5% AcOH and 25% AcOH it is 1g/1.9L and 1g/1.1L respectively. **IRRITANT.** [*Beilstein 2* H 515, *2* IV 1826.]

**Strontium salicylate** [526-26-1] **M 224.7.** It crystallises from hot water (4ml/g) or EtOH. [*Beilstein 10* IV 125.]

**Strontium tartrate** [868-19-9] **M 237.7.** It crystallises from hot water. [*Beilstein 3* IV 1219.]

**Tantalum pentaethoxide** [6074-84-6] **M 406.3, b 147°/0.2mm, 202°/10mm, pK<sub>1</sub><sup>25</sup> 9.6, pK<sub>2</sub><sup>25</sup> 13.0 (for tantalic acid).** Purify it by distillation under reduced pressure. It aggregates in \*C<sub>6</sub>H<sub>6</sub>, EtOH, MeCN, pyridine and diisopropyl ether. [Bradley et al. *J Chem Soc* 726 1955, Bradley et al. *J Chem Soc* 5 1956, *Beilstein 1* IV 1312.]

**Tetraallyltin (tetraallylstannane)** [7393-43-3] **M 283.0, b 52°/0.2mm, 69-70°/15mm, d<sub>4</sub><sup>20</sup> 1.179, n<sub>D</sub><sup>20</sup> 1.536.** Possible contaminants are allyl chloride and allyltin chloride. Check the <sup>1</sup>H NMR and IR [Fishwick & Wallbridge *J Organomet Chem* 25 69 1970], and if impure, dissolve it in Et<sub>2</sub>O and shake it with a 5% aqueous solution of NaF which precipitates allyltin fluoride. Separate the Et<sub>2</sub>O layer, dry (MgSO<sub>4</sub>), and distil it at ~0.2mm. It decomposes slightly on repeated distillation. [O'Brien et al. *Inorg Synth XIII* 75 1972, Fishwick et al. *J Chem Soc (A)* 57 1971, *Beilstein 4* III 1922.]

**Tetrabutylammonium borohydride** [33725-74-5] **M 257.3, m 128-129°.** Purify it by recrystallisation from EtOAc followed by careful drying under vacuum at 50-60°. Samples purified in this way showed no signs

of loss of *active* H after storage at room temperature for more than 1 year. Nevertheless samples should be stored at *ca* 6° in tightly stoppered bottles if kept for long periods. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>. [Raber & Guida *J Org Chem* **41** 690 1976, Brändström et al. *Tetrahedron Lett* 3173 1972.]

**Tetrabutyl orthotitanate monomer (titanium tetra-*n*-butoxide)** [5593-70-4] **M 340.4, b 142°/0.1mm, 134-136°/0.5mm, 160°/0.8mm, 174°/6mm, 189°/13mm, d<sub>4</sub><sup>35</sup> 0.993, n<sub>D</sub><sup>25</sup> 1.49.** Dissolve it in \*C<sub>6</sub>H<sub>6</sub>, filter if solid is present, evaporate and vacuum fractionate through a Widmer 24 inch column. The ester hydrolyses when exposed to air to give hydrated ortho-titanic acid. The titanium content can be determined thus: weigh a sample (*ca* 0.25g) into a weighed crucible and cover it with 10ml of H<sub>2</sub>O and a few drops of conc HNO<sub>3</sub>. Heat (hot plate) carefully till most of the H<sub>2</sub>O has evaporated. Cool and add more H<sub>2</sub>O (10ml) and conc HNO<sub>3</sub> (2ml), and evaporate carefully (no spillage) to dryness and ignite the residue at 600-650°/1 hour. Weigh the residual TiO<sub>2</sub>. [Bradley et al. *J Chem Soc* 2773 1952, Speer *J Org Chem* **14** 655 1949, *Beilstein* 1 II 398, 1 III 1515, 1 IV 1415.]

**Tetrabutyl tin (tin tetrabutyl)** [1461-25-2] **M 347.2, m -97°, b 94.5-96°/0.28mm, 145°/11mm, 245-247°/atm, d<sub>4</sub><sup>20</sup> 1.0559, n<sub>D</sub><sup>24</sup> 1.473.** Dissolve it in Et<sub>2</sub>O, dry it over MgSO<sub>4</sub>, filter, evaporate and distil it under reduced pressure. Although it does not crystallise easily, once the melt has crystallised, then it will recrystallise more easily. It is soluble in Et<sub>2</sub>O, Me<sub>2</sub>CO, EtOAc and EtOH, but insoluble in MeOH and H<sub>2</sub>O, and shows no apparent reaction with H<sub>2</sub>O. [Johnson & Fritz *J Org Chem* **19** 74 1954, Staveley *J Chem Soc* 1992 1954, Van der Kerk & Luitzen *Org Synth Coll Vol* **IV** 822 1963, *Beilstein* 4 III 1920, 4 IV 4312.]

**Tetraethyl lead** [78-00-2] **M 323.5, m -38°, b 84-85°/15mm, 200°, 227.7°(dec), d<sub>4</sub><sup>20</sup> 1.653, n<sub>D</sub><sup>20</sup> 1.5198.** Its more volatile contaminants can be removed by exposure to a low pressure (by continuous pumping) for 1 hour at 0°. Purify it by stirring with an equal volume of H<sub>2</sub>SO<sub>4</sub> (d 1.40), keeping the temperature below 30°, repeating this process until the acid layer is colourless. It is then washed with dilute Na<sub>2</sub>CO<sub>3</sub> and distilled water, dried with CaCl<sub>2</sub> and fractionally distilled at low pressure under H<sub>2</sub> or N<sub>2</sub> [Calingaert *Chem Rev* **2** 43 1926]. It prevents "knocking" in petrol combustion engines. [Milde & Beatty *Adv Chem Res* **23** 306-318 1959, *Beilstein* 4 H 639.] **VERY POISONOUS.**

**Tetraisopropyl orthotitanate (titanium tetraisopropoxide)** [546-68-9] **M 284.3, m 18.5°, 18-20°, b 80°/2mm, 78°/12mm, 228-229°/755mm.** Dissolve it in dry \*C<sub>6</sub>H<sub>6</sub>, filter if a solid separates, evaporate and fractionate. It is hydrolysed by H<sub>2</sub>O to give solid Ti<sub>2</sub>O(*iso*-OPr)<sub>2</sub> **m ca 48°.** [Bradley et al. *J Chem Soc* 2027, 1952, Bradley et al. *J Chem Soc* 469 1957, *Beilstein* 1 II 328, 1 IV 1469S.]

**Tetrakis(diethylamino) titanium [(titanium tetrakis(diethylamide))]** [4419-47-0] **M 336.4, b 85-90°/0.1mm, 112°/0.1mm, d<sub>4</sub><sup>30</sup> 0.93, n<sub>D</sub><sup>30</sup> 1.54.** Dissolve it in \*C<sub>6</sub>H<sub>6</sub>, filter if a solid separates, evaporate under reduced pressure and distil it. It is an orange liquid which reacts violently with alcohols. [Bradley et al. *J Chem Soc* 3857 1960, *Beilstein* 4 IV 313.]

**Tetramethyloxorhenium [(CH<sub>3</sub>)<sub>4</sub>ReO]** [53022-7-1] **M 262.4, m ~45° (no dec <150°).** It is best prepared by adding a 1M solution of MeLi in Et<sub>2</sub>O (80ml, 80mmol) slowly to a stirred suspension of ReOCl<sub>3</sub>(PPh<sub>3</sub>)<sub>2</sub> (8.32g, 10mmol, cf [17442-18-1]) in Et<sub>2</sub>O (120ml) at -78°, and the mixture is allowed to warm slowly to ~25° to form a dark brown solution which is stirred further for 0.5 hours. The excess of MeLi is destroyed by dropwise addition, at -78°, of H<sub>2</sub>O (~10ml), and allowed to warm to ~25°, and then cooled again, to -30°; and H<sub>2</sub>O<sub>2</sub> (2.5g, 30% diluted in 20ml of H<sub>2</sub>O) is added slowly dropwise with vigorous stirring. The colour of the solution changes to red rapidly and stirring is continued for 0.5 hours after the temperature has risen to ~25°. This mixture is then cooled to -78°, the Et<sub>2</sub>O layer is filtered free from ice, dried first with CaCl<sub>2</sub> at ~25° then with molecular sieves, cooled again to -78°, filtered, and the solvent is carefully evaporated at -50°. The red-purple (carmine) crystalline residue is sublimed onto a probe at -78°, in a 1mm vacuum, to give Me<sub>4</sub>ReO (1.26g, ~48%; but can be as high as 70% by very careful removal of solvent). Lower yields are obtained if due care is not taken during the removal of solvent *in vacuo*. The low melting solid is extremely volatile subliming readily even at -35°/10<sup>-3</sup>mm; it co-distils at low temperatures with Et<sub>2</sub>O or petroleum ether, and always occurs during removal of solvents (see earlier), hence the temperature has to be kept as low as possible, and a cold probe should be

included in the system to avoid losses. In an alternative method  $\text{ReOCl}_4$  (3mmol) is reacted with  $\text{MeLi}$  (12mmol) in  $\text{Et}_2\text{O}$  at  $-78^\circ$  to give a 20% yield of  $\text{Me}_4\text{ReO}$ , although this low yield may well be due to evaporative losses.

$\text{Me}_4\text{ReO}$  is thermally stable and can be stored indefinitely *in vacuo*. It is very soluble in petroleum ethers,  $\text{Et}_2\text{O}$ ,  $\text{CS}_2$ ,  $\text{CCl}_4$ , but these solutions slowly decompose with time even in the absence of air. However, solutions in THF and aromatic hydrocarbons are more stable. It is surprisingly unreactive at  $\sim 25^\circ$  in  $\text{Et}_2\text{O}$ ,  $\text{H}_2\text{O}$ , mineral acids, alcohols,  $\text{H}_2$ ,  $\text{CO}$ ,  $\text{CO}_2$ ,  $\text{SO}_2$ ,  $\text{HCl}$  gas, butadiene, tertiary phosphines, amines, tetramethylthiuram disulfide and 1,1-diphenyl-2-picrylhydrazyl. It is however, highly reactive with air and oxygen producing instantly white fumes and crystals of  $\text{MeReO}_3$  (methyltrioxorhenium, MTO, the carrier catalyst, see [70197-13-6]). In  $^*\text{C}_6\text{H}_6$  solution with air, it slowly oxidises to form yellow-orange products (cf yellow *cis*- $\text{Me}_3\text{ReO}_3$  [56090-01-8]). In the mass spectrum, ions of Re isotopes are identified ( $^{185}\text{Re}$  37%,  $^{187}\text{Re}$  63%) with the natural isotopic abundance. The solid state IR (at  $\sim -250^\circ$  co-condensed with argon) has  $\nu_{\text{max}}$  at 2974s(CH str), 1231s(CH def), 1018s(Re=O str), 1004sh,br, 746m(Me rock), 552m and 529s(ReC str)  $\text{cm}^{-1}$ , and IR (de-oxygenated  $\text{CS}_2$ ) has  $\nu_{\text{max}}$  at 2980s, 1370m, 1002s, 749w, 520m  $\text{cm}^{-1}$ . [Mertis, Williamson and Wilkinson *JCS, Dalton Trans* 607 1975, Mertis & Wilkinson *JCS, Dalton Trans* 1488 1976, Beattie & Jones *Inorg Chem* 18 2318 1979.]

**Tetramethyltin (tetramethylstannane, tin tetramethyl)** [594-27-4] **M 178.8, m  $-55^\circ$ , b 76.6 $^\circ$ /748mm, 74-75 $^\circ$ /atm, 78 $^\circ$ /atm,  $d^{25}$  1.291,  $n_D^{20}$  1.441.** Organotin compounds are **HIGHLY TOXIC**, and all due precautions should be taken when handling them.  $\text{Me}_4\text{Sn}$  is an ethereal smelling liquid that is highly flammable with a flash point of less than  $20^\circ$ , and should be stored at low temperatures. An improved preparation of  $\text{Me}_4\text{Sn}$  has been achieved by carrying out the reaction in the higher boiling di-*n*-butyl ether, and preparing the Grignard reagent *in situ*. Under strictly anhydrous conditions, to Mg turnings (50g, 2.06g.atoms) in dry *n*- $\text{Bu}_2\text{O}$  (600ml, containing a few crystals of  $\text{I}_2$ ) is added dropwise a solution of  $\text{MeI}$  (225g, 1.59moles, freshly distilled) in *n*- $\text{Bu}_2\text{O}$  (600ml) with gentle stirring. After the initial reaction has started, the  $\text{MeI}$  solution is added at such a rate as to keep the solution refluxing gently, requiring *ca* 3 hours. Cool the solution to  $\sim 25^\circ$  and add to it, dropwise, anhydrous  $\text{SnCl}_4$  (75g, 0.29 moles) while sustaining gentle reflux (*ca* 2-2.5 hours). The mixture is then refluxed steadily ( $85-95^\circ$ ) for 1 hour, and set aside at  $\sim 25^\circ$  for several hours. The reflux condenser is replaced by a Claisen head and the mixture is distilled; whereby  $\text{Me}_4\text{Sn}$  and *n*- $\text{Bu}_2\text{O}$  distil together at  $\sim 85-95^\circ/\text{atm}$ . It dissolves stopcock silicone greases. The  $\text{Me}_4\text{Sn}$  is purified by fractionation through a Todd column (35-40 plates) to give pure *stannane* with **b 76.6 $^\circ$ /748mm** in 91% yield. [Edgell & Ward *J Am Chem Soc* 76 1169 1954.] Its IR (liquid film) has  $\nu_{\text{max}}$  at 2920 (vs) and 3000 (vs) C-H str, 1443 (s) and 1198 (s)  $\text{CH}_3$  deformation, 770 (vs)  $\text{CH}_3$  rocking, **528 (s) C-Sn str**  $\text{cm}^{-1}$ , and from Raman spectrum 160 (vs) and 145 (vs)  $\text{cm}^{-1}$  for **C-Sn-C** bending [Taimsalu & Wood *Trans Farad Soc* 59 1754 1963, Edgell & Ward *J Am Chem Soc* 77 6486 1955]. The  $^{119}\text{Sn}$  NMR (15.5 MHz, at  $28^\circ\text{C}$ , 12mm tube not spinning, neat  $\text{SnCl}_4$  as external standard) has  $\delta$  at -147.8 ( $J_{\text{H-}^{119}\text{Sn}} = 54.0\text{Hz}$  from 60MHz proton spectrum) [Lassigne & Wells *Can J Chem* 55 927 1977, for  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra see Singh *J Organomet Chem* 99 251 1975.] [Beilstein 4 H 631, 4 I 583, 4 II 1010, 4 III 1917, 4 IV 4307.]

**Tetraphenyltin** [595-90-4] **M 427.1, m 224-225 $^\circ$ , 226 $^\circ$ .**  $\text{SnPh}_4$  forms yellow crystals from  $\text{CHCl}_3$ , petroleum ether (b 77-120 $^\circ$ ), xylene or \*benzene/cyclohexane, and is dried at  $75^\circ/20\text{mm}$ . [Gilman & Rosenberg *J Am Chem Soc* 74 531 1952, Beilstein 16 IV 1592.]

**Tetra-*n*-propylammonium perruthenate (TPAP, tetrapropyl tetraoxoruthenate)** [114615-82-6] **M 351.4, m 160 $^\circ$ (dec).** It is a dark green solid that is stable at  $\sim 25^\circ$  for long periods without much deterioration, and is best kept in a refrigerator. It is a strong oxidant, and should not be heated directly as it may explode. Small amounts decompose in a flame at  $\sim 155^\circ$  in air. It can be washed with aqueous *n*-propanol, then  $\text{H}_2\text{O}$  and dried over  $\text{KOH}$  in a vacuum. It is soluble in  $\text{CH}_2\text{Cl}_2$  and  $\text{MeCN}$ . Generally, at ambient temperatures in these solvents it is most useful, when in conjunction with *N*-methylmorpholine *N*-oxide, for the oxidation of primary and secondary alcohols, including allylic and benzylic alcohols, lactol oxidation, heteroatom oxidation, and some cleavage reactions. [Dengel et al. *Transition Met Chem* 10 98 1985, Griffith et al. *J Chem Soc, Chem Commun* 1625 1987, for a review see Ley et al. *Synthesis* 639 1994, and for its catalytic activity see chapter on "Catalysis".]

§ Polymer supported reagent is available commercially.

**Thallium (I) acetate (TIOAc)** [563-68-8] M 263.4, m 131°. Thallous (I) acetate crystallises from EtOH (needles) or Me<sub>2</sub>CO. The salt forms white deliquescent crystals (specific gravity 3.68) which are soluble in H<sub>2</sub>O and EtOH. Store it in sealed containers. It forms TIOAc·AcOH crystals which melt at 64°. [Beilstein 2 H 115, 2 II 119, 2 III 195.] **POISONOUS — HANDLE CAREFULLY WITH GLOVES, and work in a well-ventilated fume hood.**

**Thallium (III) triacetate [Tl(OAc)<sub>3</sub>]** [2570-63-0] M 381.5, m 182°(dec.). The salt is prepared by stirring thallic oxide (50g) in AcOH (300ml) at 65° until all the brown-black solid has dissolved (~24 hours), then filter and cool to obtain the salt as white crystals which are collected by decantation and dried over P<sub>2</sub>O<sub>5</sub> *in vacuo* as they turn brown in air. A second crop can be obtained by concentrating the filtrate (total yield 73g, 99%). On a large scale (457g) of the brown trioxide was treated with AcOH (2L) containing Ac<sub>2</sub>O (110ml) by stirring at 80-90°, the walls of the flask were washed down with a further volume of AcOH (500ml). Heating (90-100°) and stirring was continued for 2 hours (all the brown oxide had dissolved), the hot solution was filtered using a sintered glass funnel, cooled to ~20°, and collected onto a glass frit. The solid was dissolved in the minimum volume of AcOH (~800ml) at 80-90°, filtered through a glass frit and allowed to cool first to ~20° then to ~15°, and collected by filtration. The Tl(OAc)<sub>3</sub> was dried by spreading it in a large dish in a desiccator over KOH pellets overnight to give ~97% pure salt (467g, 61% yield). It promotes the quantitative photochemical decarboxylation of a variety of aliphatic acids to yield a high proportion of the corresponding alkanes with some alkenes, dialkyls and other minor products at wavelength 3500Å; whereas at wavelength 2537Å the alkyldimers are formed in high yields. [Kochi & Bethea *J Org Chem* 33 75 1968.] [Beilstein 2 H 115, 2 III 195.] **POISONOUS — HANDLE CAREFULLY WITH GLOVES, and work in a well-ventilated fume hood.**

**Thallium (III) trifluoroacetate [Tl(tfl)<sub>3</sub>, Tl(OCOCF<sub>3</sub>)<sub>3</sub>]** [23586-53-0] M 543.4, m dec slowly >100°, 213° (dec.). This thallation reagent is prepared in 90-100% yield by heating a suspension of thallium(III) oxide in TFA containing 10-20% of H<sub>2</sub>O until clear. This colourless solution can be used directly in thallation reactions or a granular solid salt can be isolated by evaporation *in vacuo*, and then used in inert solvents such as MeCN. The salt does not have a sharp melting point and decomposes slowly >100°. It is water sensitive but can be stored at ~25° in a stoppered container away from light without appreciable decomposition for long periods. A solution of the salt in TFA, however, can tolerate up to ~20% of H<sub>2</sub>O to give a 0.8M solution before serious hydrolysis of the salt occurs. It thallated aromatic rings (e.g. PhCl, PhF, toluene, xylene, *o*-benzoic acid) in TFA to produce the corresponding ArylTl(tfl)<sub>2</sub> at low temperatures which can be isolated in 70-100% yields, or reacted further (e.g. with KI) to displace the Tl(tfl)<sub>2</sub> group and form the respective iodides. [McKillop et al. *J Am Chem Soc* 93 4841 1971, Taylor et al. *Org Synth Coll Vol VI* 709 1988, see Fieser & Fieser's *Organic Reagents* for further applications, Beilstein 2 II 186.] **POISONOUS — HANDLE CAREFULLY WITH GLOVES, and work in a well-ventilated fume hood.**

**Thallic (I) ethoxide (TIOEt)** [20398-06-5] M 249.4, d<sub>4</sub><sup>25</sup> 3.522, n<sub>D</sub><sup>20</sup> 1.676. The alkoxide is prepared by refluxing dry EtOH in a modified Soxhlet extractor whereby the hot alcohol leaches thallium shot or turnings (see below) placed in the vapours of refluxing alcohol while oxygen flows through the apparatus which is protected from moisture with a soda lime drying tube at the top of the condenser. The apparatus devised by Fieser and Fieser can be used. Thallium shot (m 303°) is made by holding a clean piece of metal with tongs and heated with a blow torch allowing the molten metal to drip in ~2L of cold H<sub>2</sub>O. *Alternatively*, thallium turnings made with a pencil sharpener from thallium rods can be used. A Soxhlet flask (500ml) containing dry EtOH (300ml) is placed under Tl shot (100ml, dried by pressing with filter paper) in the crucible above, and the EtOH is refluxed while a rapid stream of dry O<sub>2</sub> is made to flow just beneath the metal. The volume of boiling EtOH in the flask is maintained at ~300ml by further addition of dry EtOH. When all the metal has been converted to TIOEt (12-16 hours), and all the oily alkoxide has run into the flask which will form a separate heavy oily liquid layer (note that TIOEt had d = 3.522), boiling is stopped and the volume of the alkoxide solution is adjusted to 300ml which would give a saturated solution containing 9g of TIOEt/100ml. Store away from moisture. *Alternatively*, the oily layer can be sucked out with a vacuum into a container, but great care should be taken to avoid contact with moisture. Thallium in the ethoxide solution can be determined as TIOH by cooling it in an ice bath without separating the excess of EtOH (to avoid partial conversion of TIOH to black Tl<sub>2</sub>OH), and adding an equal volume of boiled and re-cooled H<sub>2</sub>O (i.e. free from O<sub>2</sub>) and evaporating *in vacuo* to give yellow

adding an equal volume of boiled and re-cooled H<sub>2</sub>O (i.e. free from O<sub>2</sub>) and evaporating *in vacuo* to give yellow crystalline TIOH. [Fieser & Fieser's *Reagents for Organic Synthesis* 2 409 1969, Dönges in *Handbook of Preparative Inorganic Chemistry* (Ed. Brauer) Academic Press Vol I p 877-878 1963, Freudenberg & Uthemann *Chem Ber* 52 1508 1919.] **POISONOUS—handle with gloves, and work in a good fume hood.**

**Tin (II) acetate (stannous acetate)** [638-39-1] **M 236.8, m 180-182°, 182.55-183°, b 239-241° (some dec under N<sub>2</sub>), d 2.31.** It is prepared by dissolving blue-black SnO<sub>2</sub> (25g) in refluxing 50%v/v AcOH (200ml), evaporating *in vacuo* (over KOH) and subliming the white residue *in vacuo* at 150-155° (~96% yield of white orthorhombic crystals). *Alternatively*, finely divided Sn is refluxed in glacial AcOH for 80-90 hours [not less as Sn(OAc)<sub>2</sub>·2AcOH is formed] and isolated as before. It hydrolyses slowly in H<sub>2</sub>O to give blue Black SnO<sub>2</sub>, Me<sub>2</sub>CO, CO<sub>2</sub> and H<sub>2</sub>, can be stored for long periods under N<sub>2</sub>, and is soluble in EtOH (2.9%), Me<sub>2</sub>CO (1.4%) and 2N AcOH (33%). It is a useful reducing (and acetylating) agent. [Donaldson et al. *J Chem Soc* 5942 1964.]

**Titanium (IV) methoxide** [992-92-7] **M 172.0, m 200-210°, b 243°/52mm.** It is extremely sensitive to moisture. Dissolve it in H<sub>2</sub>O-free \*C<sub>6</sub>H<sub>6</sub>, filter, evaporate and distil it *in vacuo* under N<sub>2</sub>. It is **FLAMMABLE** and **TOXIC**. [Bradley et al. *Metal Alkoxides* Academic Press 1978, ISBN 0121242501.]

**Titanocene dichloride** [1271-19-8] **M 248.9, m 260-280°(dec), 289.2°, 298-291°, d<sub>4</sub><sup>20</sup> 1.60.** It forms bright red crystals from toluene or xylene/CHCl<sub>3</sub> (1:1) and sublimates at 190°/2mm. It is moderately soluble in EtOH and insoluble in Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub>, CS<sub>2</sub>, CCl<sub>4</sub>, petroleum ether and H<sub>2</sub>O. The crystalline *dipicrate* explodes on melting at 139-140°. [Wilkinson et al. *J Am Chem Soc* 75 1011 1953, IR: Wilkinson & Birmingham *J Am Chem Soc* 76 4281 1954, NMR and X-ray: Glivicky & McCowan *Can J Chem* 51 2609 1973, Clearfield et al. *J Am Chem Soc* 53 1622 1975, *Beilstein* 16 IV 1769.]

**Tri-*n*-butyl tin chloride** [1461-22-9] **M 325.5, b 98-100°/0.4mm, 140-152°/10mm, 172°/25mm, d<sub>4</sub><sup>20</sup> 1.21, n<sub>D</sub><sup>20</sup> 1.492.** Fractionate it in an inert atmosphere and seal it in small aliquots in glass ampoules. It is sensitive to moisture. [Jones et al. *J Chem Soc* 1446 1947, Kocheshkov *J Gen Chem USSR* 4 1359 1934, Kocheshkov *J Gen Chem USSR* 5 211 1935, *J Appl Chem* 6 93 1936, *Beilstein* 4 III 1926, 4 IV 4330.]

**Tributyl tin hydride** [688-73-3] **M 291.1, b 76°/0.7mm, 81°/0.9mm, d<sub>4</sub><sup>20</sup> 1.098, n<sub>D</sub><sup>20</sup> 1.473.** Dissolve it in Et<sub>2</sub>O, add quinol (500mg for 300ml, to stabilise it), dry over Na<sub>2</sub>SO<sub>4</sub>, filter, evaporate and distil it under dry N<sub>2</sub>. It is a clear liquid if dry and decomposes very slowly. In the presence of H<sub>2</sub>O, traces of tributyl tin hydroxide are formed in a few days. Store it in sealed glass ampoules in small aliquots. It is estimated by reaction with aqueous NaOH when H<sub>2</sub> is liberated. **CARE:** stored samples may be under pressure due to liberated H<sub>2</sub>. [Van Der Kerk et al. *J Appl Chem* 7 366 1937, Ono et al. *Tetrahedron* 41 4013 1985, Neuman *Synthesis* 665 1987, *Curran Synthesis* 417 1988, *Beilstein* 4 IV 4312.]

**Triethyltin hydroxide** [994-32-1] **M 222.9, m 49-50°, b 153-155°/20mm.** Treat it with HCl, followed by KOH, and filter it to remove diethyltin oxide [Prince *J Chem Soc* 1783 1959]. [*Beilstein* 4 H 633, 4 I 585, 4 II 1012, 4 III 1924, 4 IV 4325.]

**Trimethyloxonium tetrafluoroborate** [420-37-1] **M 147.9, m 141-143°(sinters, open capillary), 179.6-180.0°(dec), 210-220°(dec).** The salt must be a white crystalline solid **m** ~179.6-180.0° (dec, sealed tube). Under a N<sub>2</sub> atmosphere (e.g. Dry Box), wash it twice with CH<sub>2</sub>Cl<sub>2</sub>, then twice with Na-dried Et<sub>2</sub>O, and dry by passing dry N<sub>2</sub> over the salt until free from Et<sub>2</sub>O [Curphey *Org Synth Coll Vol VI* 1019 1988]. The oxonium salt, purified in this way, can be handled in air for short periods. The sample kept in a desiccator (Drierite) for 1 month at -20° had an unaltered melting point, and samples stored in this way for >1 year are satisfactory for alkylations. <sup>1</sup>H NMR in liquid SO<sub>2</sub> in a sealed tube had a single peak at δ 4.54 (impurities have δ at 3.39). [Meerwein *Org Synth Coll Vol V* 1096 1973.] If the sample looks good, dry it in a vacuum desiccator for 2 hours (25°/1mm) and store it under N<sub>2</sub> at -20°. The melting point depends on heating rate. [*Beilstein* 1 IV 1248.]

**Cis-Trimethylrhenium dioxide** [*cis*-trimethyldioxorhenium(VII), *cis*-(CH<sub>3</sub>)<sub>3</sub>ReO<sub>2</sub>] [56090-01-8] **M 263.3, m 10-11°.** This yellow dioxide is best prepared, in ~60% yield, from Me<sub>4</sub>ReO (preceding oxide [53022-7-1],

e.g. 0.5g) in Et<sub>2</sub>O (e.g. 40ml) at -78° by exposing the solution to nitric oxide when the red colour disappears while allowing to warm to ~25° with stirring. After 1 hour (without isolating the nitroso intermediate), the Et<sub>2</sub>O is removed at -40° *in vacuo* and the yellow crystalline residue is sublimed at 10<sup>-3</sup>mm onto a probe at -78° to give *cis*-(CH<sub>3</sub>)<sub>3</sub>ReO<sub>2</sub>, which can be recrystallised from petroleum ether solution on cooling to -78°, or distilled *in vacuo*. It is reasonably stable at room temperature, and like the starting material it fumes in air forming needles of MTO, the carrier catalyst (see above). Store it under N<sub>2</sub> or argon in a sealed ampoule. Its IR spectrum is consistent with the trigonal bipyramidal structure, where bands due to methyl vibrations are present together with two strong bands at  $\nu_{\max}$  992 and 951 cm<sup>-1</sup> attributable to *cis*-ReO<sub>2</sub>. This structure is also consistent with the <sup>1</sup>H NMR in CS<sub>2</sub> or deuteriotoluene which has two sharp resonances at  $\delta$  2.07 and 2.50 in the ratio 1:2 in which the higher field band is assumed to be from the axial methyl group. The lines do not broaden on cooling to -78° or heating to 85° in toluene solution which is consistent with a rigid molecule. The MS is as predicted from the two isotopes of Re (see above). [Mertis & Wilkinson *JCS, Dalton Trans* 1488 1976, Beattie & Jones *Inorg Chem* 18 2318 1979.]

**Trimethyltin chloride (chlorotrimethylstannane)** [1066-45-1] **M 199.3, m 37.5-39.5°, 42°, b 45-47°/10mm, 152°/760mm, 154-156°/atm.** Me<sub>3</sub>SnCl forms colourless needles that have **HIGHLY TOXIC** vapours and is best purified by distillation at atmospheric pressure or in a vacuum. It has been prepared by dropwise addition of a solution of Et<sub>2</sub>O saturated with dry HCl (20ml) to freshly distilled (dimethylamino)trimethylstannane (2.21g, b 126°/atm,  $d_{25}^{25}$  1.274,  $n_D^{20}$  1.463 [993-50-0]) in dry Et<sub>2</sub>O (10ml). Me<sub>2</sub>NH.HCl precipitated out, was filtered, the solvent was evaporated at <20°/15mm, and the residue was distilled at 760mm providing the colourless crystalline solid Me<sub>3</sub>SnCl (1.80g, 89.7%). It sublimes slowly at ~25°/20mm. [Jones & Lappert *J Organomet Chem* 3 295 1965, for <sup>1</sup>H and <sup>13</sup>C NMR see Singh *J Organomet Chem* 99 251 1975.] It is also available commercially as 1.0M solutions in hexanes or THF. Like the reaction of Me<sub>2</sub>N-SnMe<sub>3</sub>, the Me<sub>3</sub>SnCl reacts with H<sub>2</sub>O on heating to give Me<sub>3</sub>SnOH (almost quantitatively, [56-24-6]) as white crystals **m 118°**, which sublime at their melting point. It has IR (Nujol mull) with  $\nu_{\max}$  at 3620 (m), 2985 (m), 2915 (m), 1405 (w), 1195 (m), 920 (s), 765 (s), **540** (s) **Sn-C** str, 370 (s) and 316 (w) cm<sup>-1</sup>; it exists as a *dimer* [Me<sub>3</sub>SnOH]<sub>2</sub> in the solid state (IR mull) and in \*C<sub>6</sub>H<sub>6</sub>, CCl<sub>4</sub> and in CHCl<sub>3</sub> solutions (ebullioscopy) [Okawara & Yasuda *J Organomet Chem* 1 356 1964].

**Triphenylantimony** [603-36-1] **M 353.1, m 52-54°.** Recrystallise Ph<sub>3</sub>Sb from acetonitrile [Hayes et al. *J Am Chem Soc* 107 1346 1985]. [Beilstein 16 H 891, 16 IV 1198.]

**Triphenylbismuth (bismuth triphenyl)** [603-33-8] **M 440.3, m 75-76°, 77-78°, 78.5°, b 124°/7mm,  $d_4^{98.5}$  1.6427(melt).** Dissolve Ph<sub>3</sub>Bi in EtOH, precipitate it with H<sub>2</sub>O, extract with Et<sub>2</sub>O, dry and evaporate till the residue crystallises. It has been recrystallised from EtOH and Et<sub>2</sub>O/EtOH and is a stable compound. [Forward et al. *J Chem Soc suppl* p121 1949, Pfeiffer et al. *Chem Ber* 37 4620 1904, Gilman & Yablunsky *J Am Chem Soc* 62 665 1940, UV: Jaffé *J Chem Phys* 22 1430 1954, Beilstein 4 III 1841.]

**Triphenyltin chloride (chlorotriphenylstannane, triphenylchlorostannane)** [639-58-7] **M 385.5, m 103-106°(dec), 108°(dec), b 240°/13.5mm.** Purify it by distillation, followed by recrystallisation from MeOH by adding petroleum ether (b 30-60°), **m 105-106°** [Kozeschow et al. *Chem Ber* 67 1348 1934], or by crystallisation from Et<sub>2</sub>O, or 5 parts of EtOH and a small volume of petroleum ether. [Krause *Chem Ber* 51 914 1918.] It sublimes in a vacuum. [Beilstein 16 H 914, 16 I 540, 16 II 625, 16 III 1240, 16 IV 1606.] **HIGHLY TOXIC.**

**Triphenyltin hydroxide** [76-87-9] **M 367.0, m 122-123.5°, 124-126°.** West, Baney and Powell [*J Am Chem Soc* 82 6269 1960] purified a sample which was grossly contaminated with tetraphenyltin and diphenyltin oxide by dissolving it in EtOH, most of the impurities remaining behind as an insoluble residue. Evaporation of the EtOH extract gave the crude hydroxide which was converted to *triphenyltin chloride* (above) by grinding in a mortar under 12M HCl, then evaporating the acidic solution. The chloride, after crystallisation from EtOH, had **m 104-105°**. It was dissolved in Et<sub>2</sub>O and converted to the hydroxide by stirring with excess aqueous ammonia. The ether layer was separated, dried, and evaporated to give triphenyltin hydroxide which, after crystallisation from EtOH (or MeCN) and drying under vacuum, was in the form of white crystals (**m 119-120°**), which retained some cloudiness in the melt above 120°. The hydroxide retains water (0.1-0.5 moles of water per mole)

tenaciously. [Glidewell & Liles *Acta Cryst* (B) **34** 129 1978, *Beilstein* **16** H 914, **16** I 540, **16** II 625, **16** III 1240, **16** IV 1606.]

**Tris(2,2'-bipyridine)ruthenium(II) dichloride (6H<sub>2</sub>O)** [50525-27-4] **M 748.6**. Recrystallise it from water, then from MeOH [Ikezawa et al. *J Am Chem Soc* **108** 1589 1986].

**Tris( $\eta^5$ -2,4-cyclopentadien-1-yl)samarium III [SmCp<sub>3</sub>]** [1298-55-1] **M 345.6, m 365°**. SmCp<sub>3</sub> was prepared by stirring SmCl<sub>3</sub> with cyclopentadienyl sodium [see 4989-82-1] in THF, the solvent is removed and the residue is heated at 200-250° in a high vacuum. Crystalline SmCp<sub>3</sub> is stable up to 400° and sublimes at 220°/10<sup>-4</sup>mm (also 145°/10<sup>-3</sup>mm) to give orange or yellow-orange crystals that are soluble in THF and are hydrolysed in H<sub>2</sub>O. The IR is similar to that of SmCp<sub>2</sub>. [Wilkinson & Birmingham *J Am Chem Soc* **76** 6210 1954, Birmingham & Wilkinson *J Am Chem Soc* **78** 42 1956, Watt & Gillow *J Am Chem Soc* **91** 775 1969.] It is used in the selective reduction of specific double bonds, i.e. has regioselectivity [Qian et al. *J Organomet Chem* **344** 175 1988].

The **acetonitrile SmCp<sub>3</sub> complex** [100439-84-7] **M 386.7** forms bright yellow crystals soluble in MeCN, the **cyclohexylisocyanide SmCp<sub>3</sub> complex** [37299-04-0] is soluble in \*C<sub>6</sub>H<sub>6</sub> and has **m 148°**, and the **pyridine SmCp<sub>3</sub> complex** [114269-11-3] **M 424.7** is obtained as a yellow crystalline solid by dissolving it in pyridine and evaporating to dryness and is soluble in pyridine and MeCN. [Deacon et al. *Aust J Chem* **40** 895 1987.]

**Tris(d,d-dicampholymethanato) europium (III) [Eu(dcm)<sub>3</sub>]** [52351-64-1] **M 108.5, m 220-227.5°, 229-232°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> +28.6° (c 5.4, CCl<sub>4</sub>, and varies markedly with concentration)**. Dissolve it in pentane, filter it from any insoluble material, evaporate to dryness and dry the residue (white powder) at 100°/0.1mm for 36 hours. Its IR has  $\nu_{\max}$  at 1540cm<sup>-1</sup>. [McCreary et al. *J Am Chem Soc* **96** 1038 1974.]

**Trisodium citrate dihydrate** [68-04-2, 6132-04-3] **M 294.1, m 150°(loses H<sub>2</sub>O)**. Crystallise the salt from warm water by cooling to 0°. [*Beilstein* **3** III 1100, **3** IV 1274.]

**Tris-[(3-trifluoromethylhydroxymethylene)-d-camphorato] europium (III) [Eu(tfc)<sub>3</sub>]** [34830-11-0] **M 893.6, m 195-299° (dec), ~220°, [ $\alpha$ ]<sub>D</sub><sup>24</sup> +152° (c 2, CCl<sub>4</sub>, and varies markedly with concentration)**. Purify it by extraction with pentane, filter and the filtrate is evaporated and the residual bright yellow amorphous powder is dried at 100°/0.1mm for 36 hours. A sample purified by fractional molecular distillation at 180-200°/0.004mm gave a liquid which solidified and softened at ~130°, melted at ~180° and was analytically pure. Its IR (CCl<sub>4</sub>) has  $\nu_{\max}$  at  $\nu_{\max}$  1630-1680 cm<sup>-1</sup> and the <sup>1</sup>H NMR (CCl<sub>4</sub>) has  $\delta$  at -1.3 to 0.5 (br), -0.08 (s), 0.41 (s), 1.6-2.3 and 3.39 (s). [McCreary et al. *J Am Chem Soc* **96** 1038 1974, Goering et al. *J Am Chem Soc* **93** 5913 1971.]

**Vanadium (III) acetylacetonate (vanadyl tris[2,4-pentadienato-O,O'], V[AcAc]<sub>3</sub>)** [13476-99-8] **M 348.3, m 181-184°, 185-190°, pK<sub>1</sub><sup>25</sup> 2.92, pK<sub>2</sub><sup>25</sup> 3.5 (for aquo V<sup>3+</sup> hydrolysis)**. It crystallises from acetylacetonate in brown plates. It can be distilled in small quantities without decomposition. It is soluble in CHCl<sub>3</sub> and \*C<sub>6</sub>H<sub>6</sub>, and evaporation of a CHCl<sub>3</sub> solution yields brown crystals which are washed with cold EtOH and dried in vacuum or at 100° in a CO<sub>2</sub> atmosphere. Under moist conditions it readily oxidises [V(AcAc)<sub>3</sub> to V(AcAc)<sub>2</sub>O]. [Fernelius & Bryant *Inorg Synth* **V** 105 1957, McKaveney & Freiser *Anal Chem* **30** 526 1958, UV: *J Am Chem Soc* **80** 5686 1958, *Beilstein* **1** IV 3672.]

**Vanadyl (IV) acetylacetonate (vanadyloxa bis[2,4-pentadienato-O,O'], VO[AcAc]<sub>2</sub>)** [3153-26-2] **M 265.2, m 256-259°**. It crystallises from acetone. [cf Fernelius & Bryant *Inorg Synth* **V** 105 1957, *Beilstein* **1** IV 3672.]

**Vinylferrocene (ferrocenylethene)** [1271-51-8] **M 212.1, m 51-52.5°, b 80-85°/0.2mm**. Dissolve vinylferrocene in Et<sub>2</sub>O, wash it with H<sub>2</sub>O and brine, dry (Na<sub>2</sub>SO<sub>4</sub>), and evaporate to a small volume. Purify it through an Al<sub>2</sub>O<sub>3</sub> (Spence grade H) column by eluting the yellow band with petroleum ether (b 40-60°). The low melting orange crystals can be sublimed. The **tetracyanoethylene adduct** [49716-63-4] crystallises from \*C<sub>6</sub>H<sub>6</sub>/pentane and has **m 137-139°(dec)**. [Horspool & Sutherland *Can J Chem* **46** 3453 1968, Berger et al. *J Org Chem* **39** 377 1974, Rauch & Siegel *J Organomet Chem* **11** 317 1968, *Beilstein* **16** III 1787.]

**Vinyltributylstannane (vinyltributyltin)** [7486-35-3] M 317.1, b 104-106°/3.5mm,  $d_4^{20}$  1.081,  $n_D^{20}$  1.4751. Fractionate the stannane under reduced pressure and taking the middle fraction to remove impurities such as (*n*-Bu)<sub>3</sub>SnCl. [Seyferth & Stone *J Am Chem Soc* **79** 515 1957, *Beilstein* **16** III 1279.]

**Xylenol Orange (sodium salt)** See entry [1611-35-4] in "Heterocyclic Compounds", Chapter 4.

**Zinc acetate dihydrate** [5970-45-6] M 219.5, m 100° (loses 2H<sub>2</sub>O), 237°,  $d_4^{20}$  1.74, pK<sup>25</sup> 8.96 (for hydrolysis of Zn<sup>2+</sup> to ZnOH<sup>+</sup>). It crystallises (in poor yield) from hot water or, better, from EtOH. [*Beilstein* **2** III 193, **2** IV 114.]

**Zinc acetylacetonate** [14024-63-6, 108503-47-5 (*x* H<sub>2</sub>O)] M 263.6, m 138°. The zinc complex crystallises from hot 95% EtOH. [Fernelius & Bryant *Inorg Synth* **V** 105 1957, Wakita et al. *J Organometal Chem* **301** C17 1986, *Beilstein* **2** IV 144.]

**Zinc caprylate (zinc *n*-hexanoate)** [557-09-5] M 351.8, m 142°. It crystallises from EtOH in plates. [Howarth *J Chem Soc* 1406 1929, *Beilstein* **2** H 284.]

**Zinc formate dihydrate** [557-41-5] M 191.4, m 140°(loses H<sub>2</sub>O),  $d_4^{20}$  2.21. It crystallises from water (3ml/g). [*Beilstein* **2** H 16, **2** I 14, **2** II 21, **2** III 25, **2** IV 17.]

**Zinc *RS*-lactate trihydrate** [554-05-2, 16039-53-5 (*L*)] M 297.5. It crystallises from water (6ml/g). [*Beilstein* **3** I 107, **3** II 204, **3** III 464, **3** IV 637.]

**Zincon** (*o*-[1-(2-hydroxy-5-sulfo)-3-phenyl-5-formazono]-benzoic acid) [135-52-4] M 459.4. Main impurities are inorganic salts which can be removed by treatment with dilute acetic acid. Organic contaminants are removed by refluxing with ether. It can be recrystallised from dilute H<sub>2</sub>SO<sub>4</sub> and complexes with Zn ions (see below). [Fichter & Schiess *Chem Ber* **33** 751 1900, *Beilstein* **16** IV 421.]

**Zincon disodium salt** (*o*-[1-(2-hydroxy-5-sulfo)-3-phenyl-5-formazono]-benzoic acid di-Na salt) [135-52-4, 56484-13-0] M 484.4, m ~250-260° (dec). Zincon solution is prepared by dissolving 0.13g of the powder in aqueous N NaOH (2ml diluted to 100ml with H<sub>2</sub>O). This gives a deep red colour which is stable for one week. It is a good reagent for zinc ions but also forms stable complexes with transition metal ions. [UV-VIS: Bush & Yoe *Anal Chem* **26** 1345 1954, Hunter & Roberts *J Chem Soc* 820 1941, Platte & Marcy *Anal Chem* **31** 1226 1959] The *free acid* (see above) has been recrystallised from dilute H<sub>2</sub>SO<sub>4</sub>. [Fichter & Scheiss *Chem Ber* **33** 751 1900, *Beilstein* **16** IV 421.]

**Zinc phthalocyanine** [14320-04-8] M 580.9. Sublime it in oxygen-free N<sub>2</sub>. [*Beilstein* **26** III/IV 4257.]

**Zinc 5,10,15,20-tetraphenylporphyrin** [14074-80-7] M 678.1,  $\lambda_{\max}$  418(556)nm. Purify the porphyrin by chromatography on neutral (Grade I) alumina, followed by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/MeOH [Yamashita et al. *J Phys Chem* **91** 3055 1987].

**Zirconium (IV) acetylacetonate** [17501-44-9] M 487.7, m 190-193° (sealed capillary) 194.5-195°. 192-195°. The zirconium complex crystallises from hot 95% EtOH (m 190-191°) or petroleum ether and sublimes *in vacuo* at 140° with slight decomposition. The *decahydrate* effloresces in air and dehydrates in a 0.1mm vacuum. Its solubility in 1L at 20° is 30g in CS<sub>2</sub>, 47g in CCl<sub>4</sub>, 44g in ethylene dibromide, 56g in acetylacetonate and 200g in \*C<sub>6</sub>H<sub>6</sub>. [Young & Arch *Inorg Synth* **II** 122 1964, *Beilstein* **1** H 782, **1** II 835, **1** III 3121, **1** IV 3672.]

**Zirconium (IV) propoxide** [23519-77-9] M 327.6, b 198°/0.03mm, 208°/0.1mm,  $d_4^{20}$  1.06,  $n_D^{20}$  1.454. Although it was reported that it could not be crystallised or sublimed even at 150°/10<sup>-4</sup>mm [Bradley & Wardlaw *J Chem Soc* 280 1951], the propoxide, when properly prepared, has been purified by distillation in a high vacuum [Bradley et al. *J Chem Soc* 2025 1953]. [*Beilstein* **1** IV 1420.]

# CHAPTER 6

## CATALYSTS

### INTRODUCTION

Catalysts, and the process of catalysis, are almost as old as the chemical and physical sciences. Over the centuries a large number of metals and their derivatives, and a large variety of organic compounds, have been found to catalyse reactions successfully. The importance of catalysis in the sciences can be judged by the number of Nobel Prizes that have been awarded over the years for discoveries in this field of research. During the past decade much research emphasis has concentrated on combining metals with organic ligands, by complexation, in order to achieve greater metal solubility in a variety of solvents, not only for more efficient catalysis but also, by using chiral ligands, to achieve high, if not almost complete, regioselectivity and/or stereoselectivity. This is evidenced by the award of the 2001, 2005 and 2010 Nobel Prizes for Chemistry respectively for *Catalytic Asymmetric Synthesis* (to W. S. Knowles, R. Nyori and K. B. Sharpless), for *Catalytic Olefin Metathesis* (to Y. Chauvin, R. H. Grubbs and R. R. Schrock) and for *Catalytic Cross-Coupling Reactions* (to R. Heck, E-I Negishi and A. Suzuki). In addition to the many recently published reviews and monographs on various aspects of catalysis, a journal entitled *ChemCatChem* began in 2009, and is published by Wiley-VCH on behalf of ChemPubSoc Europe [www.chempubsoc.eu]. Original research papers, and reviews, on “cutting edge” discoveries in homogeneous, heterogeneous and enzymes catalysis are published particularly in this periodical, allowing proximity and cross communication in these fields of work.

Commercial suppliers of chemicals have availed themselves of this surge in new catalysts, and are now providing a large variety of recently discovered catalysts to the scientific community. The present chapter addresses a selection of these commercially available catalysts. Generally they can be used as supplied. However, some samples may have been kept for too long, or their shelf life may not be long, or they may have partially deteriorated; so knowledge about their purification, or assessment of purity is essential. Purification details have been provided for these, and in many cases published details of their preparations have also been included. Knowledge of the methods of preparation can provide information about probable impurities, and dictate the best purification procedures. Since several catalysts (and some ligands) have high and/or not well defined melting points, spectroscopic details have been included; such as IR, <sup>1</sup>H NMR, <sup>13</sup>C NMR and <sup>11</sup>B NMR, <sup>31</sup>P NMR (when these elements are present), optical properties if they are chiral, and occasionally MS.

This chapter is divided into two sections. The first section is devoted to heterogeneous and homogeneous catalysts. The second section contains ligands and reagents used for the preparation of these catalysts and of these ligands, as well as sections on phase transfer catalysts, and a section containing a small number of chiral auxiliaries. There are a number of substances mentioned in this chapter, or could possibly have been included here, but whose purifications are described in chapters 4, 5 and 7, that are not repeated here. However, they are referred to *via* their Chemical Abstracts Registry Numbers (CASRN), and their place in the book can be located from the CASRN index which gives their respective page entry.

*It should be pointed out that great care should be taken to ascertain that the catalyst is indeed the metal or reagent used, and not some trace of unsuspected contaminating metal. This point has been clearly made in the article by J.M. Crow entitled “When is a catalyst not a catalyst” (Chemistry World 8 46 2011). He draws attention to the papers of Buchwald & Bolm Angew Chem, Int Ed 48 5586 2009, Larsson et al. Angew Chem, Int Ed 48 5691 2009, Leadbeater & Marco Angew Chem, Int Ed 42 1407 2003, Arvela et al. J Org Chem 70 161 2005, and the measurement of trace impurities at the ppb (parts per billion) and/or ppt (parts per trillion) levels in Gonda, Tolnai and Novák Chem Eur J 16 11822 2010, and the possibility of co-catalysis.*

## CATALYSTS—PART 1

### HETEROGENEOUS METAL CATALYSTS

**Copper Chromite (Lazier catalyst, copper chromium oxide,  $2\text{CuO}\cdot\text{Cr}_2\text{O}_3$ )** [12053-18-8] **M 311.1**. It is also known as **Adkins catalyst** (with very slight difference in preparation). The catalyst is prepared by adding  $\text{Cu}(\text{NO}_3)_2\cdot 3\text{H}_2\text{O}$  (109g) to a warm solution of  $\text{Ba}(\text{NO}_3)_2$  (13g) in  $\text{H}_2\text{O}$  (140ml) at  $70^\circ$  and stirring until clear. Ammonium chromate solution, made by dissolving ammonium dichromate (63g) in  $\text{H}_2\text{O}$  (300ml) and adding concentrated aqueous ammonia (75ml, sp.gr. 0.90), is stirred in a thin stream into the solution of the above nitrates, stirred for 5 minutes further and the red-brown precipitate of copper barium ammonium chromate is filtered off, washed with  $\text{H}_2\text{O}$ , drained well and dried at  $110^\circ$ . The salt is placed in a suitable container and heated in a muffle furnace at  $350\text{--}450^\circ$  for 1 hour. Note that a spontaneous exothermic reaction occurs on heating with evolution of gas. The ignited blue-black residue ( $\sim 80\text{g}$ ), composed of copper and barium chromates and  $\text{CuO}$ , is pulverised thoroughly, suspended in aqueous 10%  $\text{AcOH}$  (600ml), stirred for 10 minutes, allowed to settle during 15 minutes, the supernatant is decanted, and the process is repeated. Filter off the solid, dry it at  $110^\circ$  and grind it in a mortar to give the catalyst ( $\sim 70\text{g}$ ) as a fine black powder. It is not affected by air or moisture and no special storing precautions are necessary. Other than the use of  $\text{Na}_2\text{Cr}_2\text{O}_7$  (Adkin) instead of  $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$  (Lazier), there seems to be very little difference in the preparation procedure or the activity of the catalyst [see Fieser & Fieser's *Reagents for Organic Synthesis* 2 82 1969]. Note that the inclusion of barium is for protection against sulfate which is a catalyst poison, and stabilises it against reduction. If a barium free catalyst is required, then it is omitted in the preparation and the amount of  $\text{Cu}(\text{NO}_3)_2\cdot 3\text{H}_2\text{O}$  should be increased to 121g without altering the preparation procedure. [Lazier & Arnold *Org Synth Coll Vol II* 142 1943, Adkins et al. *J Am Chem Soc* 72 2626 1950]. It is an effective catalyst for the reduction of aldehydes, ketones and esters to the corresponding alcohols, aliphatic but not aromatic double bonds, the reduction of amides to amines, and for the decarboxylation of carboxylic acids [see Fieser & Fieser *Reagents for Organic Synthesis* 1 186 1967]. These reactions however may have to be carried out in stainless steel bombs or autoclaves as several reactions require high temperatures ( $100\text{--}260^\circ$ ) and pressures (up to 5000psi).

**Nickel** [7440-02-0] **M 58.7, m 1455<sup>o</sup>, b 2730<sup>o</sup>, d<sup>25</sup> 8.90**. The catalytic use of the metal can be divided into (a) the Raney nickel type and (b) the non-Raney nickel type.

**Raney Nickel type:** This catalyst was introduced by the Raney Catalyst Division of Grace & Co and is made from a Ni-Al alloy ( $\sim 1:1$ , e.g. by fusion at  $1200\text{--}1500^\circ$ ). The catalyst is prepared by treating the alloy with alkali, which dissolves the Al, and the alkali is washed off as thoroughly as possible. The method was developed by Adkins and coworkers at the University of Wisconsin (USA) and the catalyst was designated as **W-1** to **W-6** (chronological, not activity, order) [Billica & Atkins *Org Synth Coll Vol III* 176 1955, and for comparison of Ws see Adkins & Krsek *J Am Chem Soc* 70 412 1948]. The catalyst has been used very extensively for reduction reactions, as well as for replacing sulfur by hydrogen in compounds; and there is a large volume of literature on this subject [see Fieser & Fieser's *Reagents for Organic Synthesis* 1 440, 723, 2 289, 293, 4 267, 5 263, 381, 570, 736, 6 502, 7 321, 8 433, 9 405, 10 339, 11 457, 12 175, 218, 338, 422, 13 265, 14 270, 16 278, 17 296]. The reactions normally require stirring or shaking at ambient temperatures, and pressures of  $\text{H}_2$  not greater than 100psi. A variety of equipment for this purpose is available commercially. The catalyst is essentially prepared by adding the Ni-Al alloy (150g) in small portions to aqueous sodium hydroxide (190g in 750ml of  $\text{H}_2\text{O}$ ) with stirring in an ice bath at  $10^\circ$  at such a rate that the temperature is kept below  $25^\circ$ , ca 2 hours. *n*-Octanol may be added to prevent excessive foaming. This reaction should be carried out in an efficient fume cupboard as  $\text{H}_2$  is liberated during the reaction (with the possibility of igniting), and it is preferable to use magnetic stirring to avoid sparking from an electric motor. When addition of the alloy is complete, allow the temperature of the mixture to rise to  $\sim 25^\circ$ , and then heat it gradually (avoiding frothing over) on a boiling water bath until evolution of  $\text{H}_2$  is complete (8-12 hours). Add distilled  $\text{H}_2\text{O}$  to restore the original volume, stir, allow the Ni to settle down, then decant the supernatant liquid. Transfer the Ni to a graduated cylinder with  $\text{H}_2\text{O}$ , allow it to settle, decant the supernatant liquid, add aqueous NaOH (25g, in 250ml  $\text{H}_2\text{O}$ ), shake the mixture thoroughly, allow the Ni to settle again and decant off the supernatant alkaline liquid. Wash the Ni with  $\text{H}_2\text{O}$  ( $\sim 750\text{ml}$  each time) by shaking, allow it to settle and decant the supernatant  $\text{H}_2\text{O}$ . Repeat the process until the washings are neutral to litmus and then a further 10 times. This may require 30 to 40 washings. Repeat the washing process three times with 95%

EtOH, then again three times with absolute EtOH. Store the Ni catalyst in bottles which are filled with absolute EtOH to the brim and stopper them. Keep them in a cold store room. Yield of Raney Ni is usually ~75g. The metal should be kept moist, or under liquid, at all times because the dry solid is highly **pyrophoric**. The more active the catalyst, the more readily it ignites. For this reason it is best to spoon the catalyst out from the settled material. When prepared as above it contains ~0.6g of Ni per ml of settled solid, and a level tablespoonful contains ~3g of Ni. [Mozingo *Org Synth Coll Vol III* 181 1955]. The activity of Raney Ni deteriorates slowly with time, but if prepared properly and stored as above it should be quite active for about 6 months. Its activity can be enhanced immensely by addition of a very small amount of triethylamine and chloroplatinic [Levering & Lieber *J Am Chem Soc* **71** 1515 1949]. It has been used both as a catalytic oxidation catalyst (with an appropriate donor), as well as a reduction catalyst [Kleiderer & Kornfeld *J Org Chem* **13** 455 1948].

**Non-Raney Nickel type:** This includes a catalyst of the composition Ni/Al<sub>2</sub>O<sub>3</sub> formed by treating Ni-Al alloy with H<sub>2</sub>O at 70° [Tyman *Chem & Ind* 404 1964], and the commercially available Ni on silica (with 60 wt% loading on Kieselguhr) and Ni on silica/alumina (with ~65 wt% loading). All Ni coordination compounds which have catalytic activity can come under this heading.

**Palladium** [7440-05-3] **M 106.4, m 1555°**, **b 3167°/atm**, **d<sub>4</sub><sup>20</sup> 0.951**, **electrical resistivity at 0° is μΩ-cm**. Pd is a silver-white metal which also occurs as a black powder or in a compressible spongy form. It is reactive towards HNO<sub>3</sub>, H<sub>2</sub>SO<sub>4</sub>, HCl and HClO<sub>3</sub>, particularly in the presence of air or oxygen; and forms dihalides with halogen at high temperatures. It has the capacity of adsorbing and absorbing large volumes of H<sub>2</sub>. It is extremely useful in chemical reactions as displayed in a limited way in this chapter. Apart from its extensive value in metal coordination chemistry it has out-classed many related metals in the field of catalysis. As well as its use in catalysis as the free metal in precipitated form, pellet or granular form, it has found extensive use in catalysis when diluted with C [7440-44-0], BaSO<sub>4</sub>, CaSO<sub>4</sub>, SrSO<sub>4</sub>, Kieselguhr, and Al<sub>2</sub>O<sub>3</sub> (see below). The following catalyst preparations have been used for many decades:

**Palladium on charcoal (5% Pd)**— A clear solution of PdCl<sub>2</sub> (8.2g, 46mmol) in concentrated HCl (20ml, 240mmol) and H<sub>2</sub>O (50ml), obtained by heating on a steam bath for 2 hours or until complete dissolution, is added with stirring to a suspension of charcoal (93g, HNO<sub>3</sub>—washed)<sup>†</sup> in H<sub>2</sub>O (1.2L) at 80° followed by 37% formaldehyde solution (80ml, 100mmol). The solution is made slightly alkaline to litmus (pH ~8.0) by careful addition of aqueous 30% sodium hydroxide with stirring which is continued further for 5 minutes longer. The catalyst is filtered off and washed with H<sub>2</sub>O (10 x 250ml), drained well then dried at 25°, first in air then over KOH or CaCl<sub>2</sub> in a desiccator. The dry catalyst (93-98g), which is in a reduced state is stored in a tightly closed vessel. [Do not dry at high temperatures as ignition may occur.]

<sup>†</sup> *Darco G-60, Norit or other carbons can be used and should be heated on a steam bath with 10% aqueous HNO<sub>3</sub> for 2-3 hours, washed free from acid (check pH of washings), and dried at 100-110° before use (see chapter on Aliphatic Compounds).*

A comparison of the rates of hydrogenation with 5% Pd/C, 5% Pt/C, 5% Rh/C and 5% Ru/C catalysts towards reduction of C=C, C=O, C-OH and benzene rings in various solvents at 25° and initial H<sub>2</sub> pressure of 1 atmosphere were made by Breitner, Roginski & Rylander [*J Org Chem* **24** 1855 1959.]

Note that palladised charcoal has also been used for **dehydrogenation**. e.g. tetralin is converted to naphthalene by boiling for 4 hours in the presence of Pd/C in a slow stream of CO<sub>2</sub>, and α-tetralone is dehydrogenated to naphthalene by boiling (1 hour, with internal temperature of 235°) in the presence of Pd/C and a slow stream of CO<sub>2</sub> in 75% yield.

**Palladium on charcoal (10% Pd)**— A clear solution of PdCl<sub>2</sub> (8.3g) in concentrated HCl (5.5ml) and H<sub>2</sub>O (40ml), prepared as above, is poured into a solution of NaOAc.3H<sub>2</sub>O (135g) in H<sub>2</sub>O (500ml) in a hydrogenation bottle, and charcoal (45g, HNO<sub>3</sub>—washed, see above<sup>†</sup>) is added and the mixture is hydrogenated until absorption ceases (1-2 hours). The catalyst is collected on a filter funnel, washed with H<sub>2</sub>O (5 x 400ml), drained well, dried in air and then in a desiccator over KOH or CaCl<sub>2</sub>; and the reduced catalyst (48-50g) is stored as above.

**Palladium chloride on charcoal (5% Pd)**— A clear solution of PdCl<sub>2</sub> (8.2g, 46mmol) in concentrated HCl (20ml, 240mmol) and H<sub>2</sub>O (50ml), obtained by heating on a steam bath for 2 hours or until complete dissolution, is diluted with H<sub>2</sub>O (140ml) and poured over charcoal (92g, HNO<sub>3</sub>—washed, see above<sup>†</sup>) in a 20cm evaporating dish, thoroughly mixed and dried, firstly on a steam bath and then in an oven at 100° with occasional stirring until the catalyst (98-100g) is completely dry and is then stored in a stoppered bottle until required.

When required, the desired aliquot of catalyst is transferred to a hydrogenation bottle (or flask) and reduced in the solvent used for hydrogenation by shaking with  $H_2$  in the hydrogenation apparatus. When absorption of  $H_2$  is complete, the catalyst is filtered off, washed with fresh solvent to remove HCl, returned to the hydrogenation bottle (which should have been rinsed with the same fresh solvent), the material for reduction is added and hydrogenation is carried out in the usual way.

**Palladium on barium sulfate (5% Pd)**— To a rapidly stirred hot ( $80^\circ$ ) solution of reagent grade  $Ba(OH)_2 \cdot 8H_2O$  (126.2g, 400mmol) in  $H_2O$  (1.2L), 6N  $H_2SO_4$  (120ml, 360mol) is added all at once then more 6N  $H_2SO_4$  is added to the  $BaSO_4$  suspension that is formed in order to be just acidic to litmus (pH  $\sim 5$ ). A clear solution of  $PdCl_2$  (8.2g, 46mmol) in concentrated HCl (20ml, 240mmol) and  $H_2O$  (50ml), obtained by heating on a steam bath for 2 hours or until complete dissolution, is subsequently added to the hot ( $80^\circ$ ) stirred  $BaSO_4$  suspension<sup>#</sup> followed by 37% formaldehyde solution (8ml, 100mmol). The suspension is made alkaline to litmus (pH  $\sim 8.0$ ) by careful addition of aqueous 30% sodium hydroxide with stirring, which is continued further for 5 minutes longer, and allowed to settle. The clear colourless supernatant liquid is decanted off, replaced by  $H_2O$ , resuspended, and washed ten times by decantation. The catalyst is collected on a 90mm medium porosity sintered glass funnel, drained until the filter cake begins to crack, washed again with  $H_2O$  (95 x 250ml), drained as completely as possible and placed in an oven at  $80^\circ$  until a dry catalyst (93-98g) is obtained. Store it as above. [This catalyst is particularly useful for Rosenmund catalytic reductions of acid chlorides (RCOCl) to aldehydes (RCHO), cf Hersberg & Cason *Org Synth Coll Vol III* 627 1955.]

<sup>#</sup>An equivalent weight of precipitated  $BaCO_3$  (93g) may be used instead for  $BaSO_4$ ; but the volume of concentrated HCl in preparing the  $PdCl_2$  solution should be reduced to 8.2ml (to avoid decomposition of the carbonate) to give **5% Pd on calcium carbonate** catalyst. [Mozingo *Org Synth Coll Vol III* 685 1955.]

**Note** that at the completion of hydrogenations with these catalysts, the used catalysts that are filtered off should be kept moist with solvent or  $H_2O$  as they may be **pyrophoric** when dry. Also, with the catalysts that have been prepared using the formaldehyde (HCHO) procedure, the spent catalysts from reduction reactions in which MeOH is used as solvent are particularly flammable, possible because MeOH/HCHO form an oxidation-reduction couple since there is no guarantee that the catalyst is completely free from HCHO.

Other commonly used supports for Pd include **CaCO<sub>3</sub>** [Busch & Stöve *Chem Ber* **49** 1063 1916], **SrCO<sub>3</sub>** [Martin & Robinson *J Chem Soc* 4911943], **Al<sub>2</sub>O<sub>3</sub>** [Alan C. Johnson US pat 2,366,409, *Chem Abstr* **39** 2001 1945], **Kieselguhr** [Rosenmud & Langer *Chem Ber* **56** 2262 1923, Sabalitschka & Moses *Chem Ber* **60** 786 1927], **silica-gel** [Fester & Brude *Chem Ber* **56** 2247 1923], **asbestos** [Zelinsky & Borisoff *Chem Ber* **57** 150 1924, Zelinsky & Turowa-Pollak *Chem Ber* **58** 1295 1925] and **polyvinyl alcohol** [Kavanagh & Nord *J Am Chem Soc* **66** 2126 1944].

**Platinum** [7440-06-4] **M 195.1, m 1772<sup>o</sup>, b 3827<sup>o</sup>, d<sup>20</sup> 21.45.** Platinum, when used directly or on a carbon, alumina or silica support, or *via* its oxide, is generally an excellent catalyst for the reduction of compounds. Hydrogenations using Pt catalysts are usually carried out at ambient temperatures, and atmospheric pressure of  $H_2$ , and the reaction can be followed by using a manometer and noting the absorption of gas. They may require, as in the reduction of aromatic rings to alicyclic rings, somewhat elevated pressures (a few to several atmospheres) and temperatures. Specialised equipment is used in such cases (see Ni catalysts above).

**Platinum Black:** It is obtained by reducing a solution of chloroplatinic acid [16941-12-1] with alkaline formaldehyde and shaking vigorously to coagulate the colloidal metal formed [Feulgen *Chem Ber* **54** 360 1921]. In addition to its use as a catalyst, it destroys  $H_2O_2$  in liquids or solvents by shaking them with the metal until liberation of  $O_2$  ceases [Cope & Cignek *Org Syn Coll Vol IV* 612 1963].

**Platinum Sponge:** Ammonium hexachloroplatinate is precipitated when recovered Pt residues are dissolved in *aqua regia* (see below), evaporated to dryness then dissolved in the minimum volume of  $H_2O$  and treated with saturated ammonium chloride. The  $(NH_4)_2PtCl_6$  is filtered off and dried at  $100^\circ$ . Ignition of this salt gives Pt sponge. [Wickers *J Am Chem Soc* **43** 1268 1921.]

**Brown & Brown Catalyst:** This is usually prepared *in situ* by adding a solution of  $NaBH_4$  in EtOH stabilised with NaOH to a suspension of charcoal (see palladium catalysts above) in a solution of chloroplatinic acid in EtOH (or aqueous EtOH) to form the supported catalyst. Excess HCl is added, followed by the substrate for reduction, then  $NaBH_4$  in EtOH stabilised with NaOH is added dropwise at such a rate as to maintain atmospheric pressure of  $H_2$ . Here the supported catalyst and  $H_2$  are generated *in situ* resulting in more effective reductions. [Brown et al. *J Org Chem* **28** 214 1963.]

**Platinum on support:** The catalyst is prepared from a cooled solution of the metal (5.0g, or  $\text{PtCl}_4$  or  $\text{H}_2\text{PtCl}_6$ ) in  $\text{H}_2\text{O}$  (50ml) and concentrated  $\text{HCl}$  (5ml) and support [11g, acid washed charcoal (see Pd above), neutral alumina, silica or asbestos (Gooch asbestos, termolyte not chrysolite, boil with concentrated  $\text{HNO}_3$ , filter, wash free from acid with  $\text{H}_2\text{O}$  and dry at  $100^\circ$ )] in a freezing mixture and treat with formalin (50ml, 40% formaldehyde). The mixture is then stirred while aqueous  $\text{KOH}$  (50g in 50ml of  $\text{H}_2\text{O}$ ) solution is added slowly and keeping the temperature below  $5^\circ$ . The temperature is then allowed to rise to  $60^\circ$  in 15 minutes. The catalyst is washed thoroughly with  $\text{H}_2\text{O}$  by decantation, finally with dilute  $\text{AcOH}$ , collected onto a suction filter and washed with  $\text{H}_2\text{O}$  until free from chloride ions and alkali, and drained well. Dry the solid at  $100^\circ$  and store it in a desiccator. This method gives a catalyst containing *ca* 30-40% of metal, but the amount can be adjusted by varying the ratio of metal to support. [Brown & Brown *J Am Chem Soc* **28** 2827, 2829 1962.]

A comparison of the rates of hydrogenation with 5% Pd/C, 5% Pt/C, 5% Rh/C and 5% Ru/C catalysts towards reduction of C=C, C=O, C-OH and benzene rings in various solvents at  $25^\circ$  and initial  $\text{H}_2$  pressure of 1 atmosphere were made by Breitner, Roginski & Rylander [*J Org Chem* **24** 1855 1959.]

**Platinum (IV) oxide (Adams catalyst,  $\text{PtO}_2$ )** [*1H<sub>2</sub>O* 12137-21-2; *xH<sub>2</sub>O* 52785-06-5] **M 227.1, M 245.1 (1  $\text{H}_2\text{O}$ ), m >450°.** Adams catalyst can be prepared in three ways.

**Firstly**, from ammonium hexachloroplatinate (IV) [16919-58-7] [3.0g,  $(\text{NH}_4)_2\text{PtCl}_6$ ] and  $\text{KNO}_3$  (30g, Analytical Grade;  $\text{NaNO}_3$  yields slightly less active catalyst) which are thoroughly mixed in a porcelain dish (**fume cupboard, use a head shield for protection**) and heated gently at first until the rapid evolution of gas subsides, then strongly until the temperature reaches  $300^\circ$  (15 minutes, no splattering). The liquid mass is heated further at  $500\text{-}530^\circ$  for 30 minutes and allowed to cool. Add  $\text{H}_2\text{O}$  (50ml) to the residue, stir and allow the brown platinum oxide ( $\text{PtO}_2$ ,  $\text{H}_2\text{O}$ ) to settle. Wash it once or twice with  $\text{H}_2\text{O}$  by decantation, collect it on a filter, wash it free from nitrates (stop washing if the solid starts to become colloidal#; traces of  $\text{NaNO}_3$  do not affect catalytic activity) and dry the oxide in a desiccator. Store it in small aliquots.

# If colloidal Pt filters through, it can be checked by testing a small volume of the filtrate for Pt by acidifying it with  $\text{HCl}$ , adding a few drops of stannous chloride when a yellow or brown colour results depending on the amount of Pt present. The yellow colour is soluble in  $\text{Et}_2\text{O}$  rendering the test more sensitive. If Pt is present in the filtrate then heat the filtrate with excess formalin and aqueous  $\text{NaOH}$ , cool, and the platinum black which separates is filtered off and worked up with other Pt residues. The weight of  $\text{PtO}_2$  obtained is about equal to half the weight of  $(\text{NH}_4)_2\text{PtCl}_6$ .

**Secondly**, purest possible chloroplatinic acid [3.5g,  $\text{H}_2\text{PtCl}_6$ ] in  $\text{H}_2\text{O}$  (10ml) is mixed with  $\text{KNO}_3$  (35g, Analytical Grade;  $\text{NaNO}_3$  yields slightly less active catalyst) in a porcelain dish, evaporate gently to dryness and heat at  $350\text{-}370^\circ$  for 10 minutes (use a strong heat source. **fume cupboard, use a head shield for protection**) when brown oxides of nitrogen evolve and brown  $\text{PtO}_2$  separates. Stir more vigorously if frothing occurs and direct an additional burner at the top of the mixture. [Take care not to allow the top of the mixture to solidify, e.g. by removing the burner that is underneath, otherwise frothing over occurs]. The temperature is raised to  $400^\circ$  in the following 15 minutes, after which liberation of fumes decreases considerably, and the temperature is raised further to  $500\text{-}550^\circ$  (*ca* 2 minutes); at this stage evolution of brown gases practically ceases but there is still a gentle evolution of gas. Maintain at this temperature (with burner at full heat) for 30 minutes by which time fusion is complete. Allow the container to cool slowly to  $\sim 20\text{-}25^\circ$ , add  $\text{H}_2\text{O}$  (50ml) and isolate the catalyst as above.

**Thirdly**, Pt metal and or Pt metal residues are dissolved in *aqua regia* [**prepared by adding slowly concentrated  $\text{HNO}_3$  (1 volume) to concentrated hydrochloric acid (3 volumes) in a glass container in an efficient fume cupboard and head protection**] using gentle heat if necessary, evaporate several times almost to dryness with concentrated  $\text{HCl}$ , dissolve the final residue in a small volume of  $\text{H}_2\text{O}$  and precipitate the metal as  $(\text{NH}_4)_2\text{PtCl}_6$  with saturated aqueous  $\text{NH}_4\text{Cl}$ . Filter off the precipitate, dry it at  $100^\circ$  and continue as in the first procedure (see above). [Adams et al. *Org Synth Coll Vol I* 463, 466 1941, Bruce *J Am Chem Soc* **58** 687 1936.]

Reactions catalysed by Adams catalyst can be carried out in alcoholic solution which is sometimes enhanced by the presence of  $\text{HCl}$ ; a variety of other organic solvents have been used which include  $\text{EtOAc}$ ,  $\text{EtOAc}+15\%$   $\text{AcOH}$ ,  $\text{EtOAc}+8\%$   $\text{EtOH}$ , and glacial  $\text{AcOH}$ ; and neat  $\text{CF}_3\text{CO}_2\text{H}$  is a particularly good solvent for the reduction of C=N in heterocyclic compounds.

**$\text{PtO}_2$  on silicic acid** has been prepared in much the same way as above (**fume cupboard and head protection**)

from chloroplatinic acid (7g) in H<sub>2</sub>O (25ml) and silicic acid (20g, 200 mesh) in a borosilicate beaker (100ml) which are stirred into a smooth paste and heated on a burner until desiccated. Heat to ~360° and add powdered NaNO<sub>3</sub> (70g) in portions while stirring and mixing vigorously with a glass rod (evolution of brown fumes occurs at each addition with frothing) and heat until gas evolution ceases. Add further molten NaNO<sub>3</sub> (20g) to the beaker with heating while scratching its sides. Cool, and transfer to a 2L beaker containing H<sub>2</sub>O (1.5L), stir for 2 hours, allow to settle, decant the supernatant and collect on a medium porosity sintered glass funnel. Wash the solid well with H<sub>2</sub>O, then with 95% EtOH (2 x) and Et<sub>2</sub>O, dry it in air for 1 hour then in a vacuum desiccator (P<sub>2</sub>O<sub>5</sub>). The dry powder is then pulverised in a mortar and sieved (200 mesh) until all the solid has passed through (use a Camel's-hair brush to help the solid through) to give the catalyst as a light brown powder (21.0g) containing ~0.14g of Pt per gram of solid. Store it in screw capped vials. It can be exposed to light and air without loss of activity. It gives very reproducible hydrogenation results and has been used analytically for the estimation of reducible groups [Vandenheuevel *Anal Chem* **28** 362 1956.]; and for the hydrogenation of methyl esters in EtOH without causing trans-esterification [Ackman & Brugher *J Lipid Res* **5** 130 1964].

**Other applications:** Prior to the development of catalysts in which Pt is coordinated to a plethora of ligands in order to achieve particular specificities (e.g. Pt with phosphine ligands below), some specificities were obtained by manipulating the metal. Examples include a **platinum-tin chloride** catalyst made from **H<sub>2</sub>PtCl<sub>6</sub> and SnCl<sub>2</sub>** for specific reduction of olefins [Cramer et al. *J Am Chem Soc* **85** 1691 1963], **PtO<sub>2</sub>/Pt black+FeCl<sub>2</sub>/Zn(OAc)<sub>2</sub>** for reduction of unsaturated aldehydes to unsaturated alcohols, e.g. cinnamaldehyde to cinnamyl alcohol [Tulley & Adams *J Am Chem Soc* **47** 3061 1925], and **PtS<sub>2</sub>/C** which is less active than the oxide but is insensitive to poisons (e.g. sulfur compounds) and will reduce a nitro group without affecting halogen substituents, e.g. chloronitrobenzenes to chloroanilines [Dovell & Greenfield *J Am Chem Soc* **87** 2767 1965].

**Rhenium metal (Re) used in preparing the following catalysts has [7440-15-5] M 186.2, m 3180°, b 5596°, d<sup>20</sup> 21.02.**

**Rhenium(II) oxide (ReO.2H<sub>2</sub>O) [12143-03-2] M 238.2.** This catalyst is prepared in a fume cupboard by the reduction of ammonium perrhenate with either Na/liqNH<sub>3</sub>, Li/liqNH<sub>3</sub>, or Li/EtNH<sub>2</sub> which give consistently ReO.2H<sub>2</sub>O. In dried apparatus used typically for Na/liqNH<sub>3</sub> reduction, liquid NH<sub>3</sub> or EtNH<sub>2</sub> (50-75ml) and NH<sub>4</sub>ReO<sub>4</sub> (1-2g, weighed) are stirred in a dry atmosphere until the salt dissolves; then Na or Li (15-30g) are added in chunks as rapidly as possible. The colourless solution becomes typically blue-black and is stirred until the colour disappears (15 minutes to 2 hours). The NH<sub>3</sub> or EtNH<sub>2</sub> are allowed to evaporate while the total volume is kept constant by addition of EtOH. The residue is transferred to a fritted glass thimble in a Soxhlet apparatus and extracted with dilute HCl (~2N) for 24 hours (the pot acquiring a brown colour), then with EtOH for an equal time. When the brown acid extract, that decolorises aqueous KMnO<sub>4</sub>, is basified with aqueous NH<sub>4</sub>OH black ReO.2H<sub>2</sub>O precipitates, and the supernatant now does not decolorise aqueous KMnO<sub>4</sub>. The solid is collected, washed with cold EtOH, and dried at 100° *in vacuo* over P<sub>2</sub>O<sub>5</sub> (conditions under which it is stable and does not dehydrate) to provide the hydrated oxide (75-85%) as a very finely divided black powder. Note that any solid left in the thimble has the same composition as this oxide. It is stored as a suspension in EtOH which tends to settle rather readily, or as a suspension in distilled H<sub>2</sub>O which requires more than 24 hours to settle out completely. Aliquots of suspensions (with or without evaporating to dryness) are used as the heterogeneous catalyst. The amount of rhenium can be determined as described by Broadbent and Selin [*J Org Chem* **28** 2343 1963]. ReO.2H<sub>2</sub>O catalyses the hydrogenation of a wide variety of organic substrates in high yields with unusual selectivity, being particularly efficient for carboxylic acids. Hydrogenations are carried out in a rocking Parr reactor under pressures of ~200 atmospheres and temperatures varying from ~20° to ~200°, and reduction times from 0.5 to ~10 hours. The catalyst prepared by the Li/liqNH<sub>3</sub> method is only marginally better than by the other methods. [Broadbent & Seegmiller *J Org Chem* **28** 2347 1963.]

**Rhenium(III) oxide (Re<sub>2</sub>O<sub>3</sub>) [12060-05-8] M 420.5.** Re<sub>2</sub>O<sub>3</sub> is prepared from a mixture of ammonium perrhenate (2.04g, 160mmol, NH<sub>4</sub>ReO<sub>4</sub>), a solution of 8N NH<sub>4</sub>Ac (20ml, 160mmol) and glacial AcOH (2ml, 35mmol) in H<sub>2</sub>O (100ml) in an ice-water bath to which is added dropwise, with vigorous stirring, an ice-cold solution of NaBH<sub>4</sub> (3.4g, 90mmol) in H<sub>2</sub>O (100ml) over a 1 hour period. After half of the NaBH<sub>4</sub> solution is added, a further amount of AcOH (2ml, 35mmol) is added in one lot. Stirring is continued for 15 minutes and the solution is centrifuged at full speed in an ultracentrifuge, the supernatant is decanted off, the oxide is washed

twice by suspension in H<sub>2</sub>O and finally re-suspended in H<sub>2</sub>O, and stored as such before use. The fine solid remains in suspension for 12-24 hours. Analysis indicated that reduction was ~85% complete giving a yield of 70-90% (~1.35g). The activity is checked by its ability to effect hydrogenation of styrene at room temperature in 6-10 hours. It catalyses a range of substrates including aliphatic olefins, maleic acid, benzene, nitrobenzene, bromonitrobenzene, naphthalene, pyridine, ketones and nitriles mostly in yields approaching 100% with (e.g. H<sub>2</sub>O or EtOH), and without solvents. Typically the substrate (0.2mol, with or without solvent) and catalyst (0.2g of Re) in a 420ml rocking Parr bomb under hydrogen pressures up to 3000psi (204 atmospheres) are heated (40-50° increments) until the pressure starts to drop. Reductions occur in the temperature range of 25-250°. [Broadbent & Johnson *J Org Chem* **27** 4400 1962.]

**Rhenium(IV) oxide (ReO<sub>2</sub>·2.5H<sub>2</sub>O)** [12036-09-8] **M 218.2**. When a vigorously stirred mixture of a solution of ammonium perrhenate (1.2g) in H<sub>2</sub>O (150ml) and an excess of granular zinc (5-10g, previously treated with a little dilute acid, to remove any surface oxide, and washed thoroughly to clean the surface) is treated dropwise over a period of several hours with 3N H<sub>2</sub>SO<sub>4</sub> (64ml) at such a rate that black finely divided rhenium oxide is suspended into the solution which is periodically decanted from the zinc; the oxide is centrifuged off, the supernatant is returned to the reaction flask and addition of acid is resumed. The acid should not be added rapidly as the finely divided rhenium oxide coats the zinc and slows (or prevents) further reaction, and does not produce a Zn-free product. At the end of the reaction all the rhenium oxide residues from centrifugation are combined, washed with H<sub>2</sub>O by decantation until the washings are free from sulfate. Conversion from perrhenate to rhenium (IV) oxide is >60%, and nearer 100% if an excess of Zn is maintained in the reaction by further addition of the metal. ReO<sub>2</sub>·2.5H<sub>2</sub>O analyses as a hydrate, is very stable, resists dehydration at 95° *in vacuo* over P<sub>2</sub>O<sub>5</sub>, and loses H<sub>2</sub>O only very slowly at 250°; however when the hydrated oxide is heated at 250° in \*C<sub>6</sub>H<sub>6</sub> under hydrogenation conditions anhydrous ReO<sub>2</sub> can be isolated from the experiment. It exists in two crystallographic forms: the α-form is obtained when synthesised below 300° (Re + 2ReO<sub>3</sub> = 3ReO<sub>2</sub>), and is irreversibly transformed into the orthorhombic β-form which is obtained when synthesised above 300° [Rogers et al. *Inorg Synth* **XIII** 142-145 1972, Rogers et al. *Inorg Synth* **XXX** 102 1995].

ReO<sub>2</sub>·2.5H<sub>2</sub>O has catalytic activity for the hydrogenation of a variety of organic substrates, and although generally less active than Nickel or palladium catalysts (i.e. requires higher temperatures and/or pressures), it is however much more effective than these for the reduction of the carboxylic acid function. It has a potentially valuable selectivity, and very few other elements exhibit this characteristic to such a degree. [Broadbent & Selin *J Org Chem* **28** 2343 1963.]

**Rhenium(VI) oxide (ReO<sub>3</sub>)** [1314-28-9] **M 234.2, b 750°, d<sup>20</sup> 6.9-7.4**. ReO<sub>3</sub> is prepared from the heptoxide (1.0g, see below) by adding it directly to *p*-dioxane (10-25ml) or tetrahydrofuran and warming gently to produce a colourless to dark green solution, and the excess solvent is removed *in vacuo* to give a sticky, tarry, black complex. The complex decomposes when carefully heated to *ca* 145° to give gaseous products and a pure, red, crystalline residue of ReO<sub>3</sub> (84-95%yield). This is finely ground and stored dry in a screw-cap vial. It forms non-hygroscopic, red/maroon-coloured crystals (with a green luster) which are air stable below 100°, and disproportionates in a vacuum to rhenium(IV) and rhenium(VII) oxides at 400°. It is inert to H<sub>2</sub>O, dilute alkali, and is thus more convenient to handle and store than the hygroscopic heptoxide which also has high catalytic activity. Analysis showed it to have the formula of the anhydrous oxide ReO<sub>3</sub>. It is oxidised by HNO<sub>3</sub> to HReO<sub>4</sub>. It has catalytic activity when used as such or when prepared *in situ* by using Re<sub>2</sub>O<sub>7</sub> and allowing it to be reduced to ReO<sub>3</sub> during the hydrogenation process. This oxide, when used directly or prepared *in situ*, is the most effective hydrogenation catalyst of all the Re catalysts. It is very efficient in reducing carboxylic and carboxamide groups without reducing the aromatic ring, with the latter producing amines. Benzoic acid, ethyl benzoate, benzaldehyde, and *m*-nitrobenzaldehyde are reduced to the respective carbinols without hydrogenolysis to the toluenes as in most catalytic hydrogenations. Like the preceding rhenium oxide catalysts hydrogenations are carried out in shaking Parr-Bomb reactors under pressures of *ca* 200 atmospheres and temperatures ranging from ~100° to ~200° depending on when the substrates begin to absorb H<sub>2</sub> at a reasonable rate. [Broadbent & Bartley *J Org Chem* **28** 2345 1963, Nechamkin & Hiskey *Inorg Synth* **III** 186 1950.] It is also obtained in quant-

itative yield from  $\text{Re}_2\text{O}_7$  (purified by sublimation) in a CO atmosphere by heating at  $175^\circ$  until the oxide is blue, the temperature is raised slowly to  $225^\circ$ , then increased to  $280^\circ$  after the colour of the oxide has turned to red [Melaven et al. *Inorg Synth* **III** 187 1950].

**Rhenium(VII) oxide (rhenium heptoxide  $\text{Re}_2\text{O}_7$ )** [1314-68-7] **M 484.4, m 300.3°, b 360.3°, d<sup>20</sup> 6.103.**

Dirhenium heptoxide is prepared by heating the metal or its lower oxides in a stream of oxygen or air at  $400^\circ$  to  $425^\circ$ . It forms canary-yellow crystals which are as deliquescent as  $\text{P}_2\text{O}_5$ , and sublimates at  $250^\circ/760\text{mm}$  and melts at  $297^\circ$  in a sealed tube [Ogawa *Bull Chem Soc Jpn* **7** 265 1932]. Great care should be taken to avoid moisture during handling and storage. It is very soluble in  $\text{H}_2\text{O}$  to form perrhenic acid  $\text{HReO}_4$ . It is also soluble in most organic solvents such as ethers, esters, alcohols, dioxane and pyridine. [Melaven et al. *Inorg Synth* **III** 188 1950.] Reduction of  $\text{Re}_2\text{O}_7$  during hydrogenation reactions generates rhenium “blacks” which mainly consist of  $\text{ReO}_3$  and is among the best catalyst for reducing a variety of carboxylic acids to alcohols (at  $\sim 160^\circ/\text{ca}$  200 atmospheres/few hours) in very good yields. The alcohols are occasionally accompanied with esters; unreduced acids rarely survive, amides are reduced to the respective amines and the usual functional groups that are reduced by Ni and Pt catalysts are also reduced with this rhenium catalyst. [Broadbent et al. *J Org Chem* **24** 1847 1959.] It is a precursor for organorhenium oxides which catalyse olefin oxidation, metathesis and other transformations. It is a *Lewis-acidic metal oxide* which forms adducts of the formula  $\text{Re}_2\text{O}_7 \cdot n\text{L}$  with a variety of O, N and S ligands such as 1,2-dimethoxyethane, 4,4'-di-*tert*-butyl-2,2'-bipyridine, and 1,4,7-trithiacyclononane respectively [Herrmann et al. *Inorg Chem* **34** 4701 1995].

**Rhenium(VII) sulfide (rhenium heptasulfide,  $\text{Re}_2\text{S}_7$ )** [12038-67-4] **M 596.9.** The heptasulfide is obtained by bubbling  $\text{H}_2\text{S}$  through a 3% solution of potassium perrhenate in 5-6 N HCl maintained at the boiling point until precipitation is complete. The suspension is allowed to stand overnight, then it is filtered through a sintered glass funnel, the solid is washed with 1:4 aqueous HCl,  $\text{H}_2\text{O}$ , dried for several days in a desiccator over  $\text{CaCl}_2$ , powdered in a mortar and stored in a common screw capped vial. The catalyst is not adversely affected by occasional exposure to air, is prepared with high catalytic reproducibility, is insoluble in concentrated non-oxidising acids, is stable under all the conditions of hydrogenation used (1000-3600psi/ $25^\circ$  to  $335^\circ/0.5$  to 3 hours), and is not reduced to form  $\text{H}_2\text{S}$ . It is remarkably resistant to “poisoning”. It catalyses the reduction of a variety of organic substrates in high yields including sulfides, with desulfurisation [e.g. diphenylsulfide to benzene (94%), and thiophenol to a mixture of benzene (60%) and cyclohexane (40%)]; and without desulfurisation [e.g. allyl phenyl sulfide to propyl phenyl sulfide (100%), and of thiophene to thiophane (70%)]. [Broadbent et al. *J Am Chem Soc* **76** 1519 1954.] When the heptasulfide is heated for 16 hours at  $350\text{-}400^\circ$  in a tube furnace under a stream of  $\text{CO}_2$  to blow out sulfur vapours, and cooled in a desiccator, it gives **rhenium disulfide** [12038-63-0]. The catalytic activity of  $\text{ReS}_2$  parallels that of  $\text{Re}_2\text{S}_7$  but to a lesser extent, and is more active than related molybdenum or cobalt catalysts. [Broadbent et al. *J Am Chem Soc* **76** 1519 1954.]

**Rhodium metal (Rh) used in preparing the following catalysts has** [[7440-16-6] **M 102.9, m 1966°, b 3727°, d<sup>20</sup> 12.41**].

**Rhodium on alumina (5% Rh loading).**  $\text{Rh-Al}_2\text{O}_3$  has been used as a general catalyst for the complete reduction of the aromatic ring in phenols, e.g. gallic acid to *cis*-3,4,5-trihydroxycyclohexane carboxylic acid [Burgstahler & Bithos *Org Synth* **42** 62 1962] requiring high pressures (2200psi) and temperatures ( $\sim 100^\circ$ ) for 8-12 hours [see also Smith & Stump *J Am Chem Soc* **83** 2739 1961, Kaye & Matthews *J Org Chem* **28** 324 1963], the hydrogenation of 1-naphthol to *cis*-1-hydroxydecalin [6psi,  $25^\circ$ , 12 hours; Meyers et al. *J Org Chem* **29** 3427 1964], resorcinol gives cyclohexane-1,3-dione [50psi,  $25^\circ$ , 16-18 hours; Sircar & Meyers *J Org Chem* **30** 3206 1965], D-mandelic acid gives pure D-hexahydromandelic acid but L-mandelic acid gives ( $\pm$ )-hexahydromandelic acid [3-4 atmospheres,  $25^\circ$ , 1.5 hours, AcOH is mandatory; Stocker *J Org Chem* **27** 2288 1962], reduction of vinylic and allylic double bonds [500psi,  $100^\circ$ , 15 minutes; Ham & Coker *J Org Chem* **29** 194 1964], reduction of oximes or aliphatic nitriles to primary amines [1-2 atmospheres,  $\sim 25^\circ$ , 1-2 hours; Freifelder et al. *J Org Chem* **27** 2209 1962, Freifelder *J Am Chem Soc* **82** 2386 1960], hydrogenolysis of ketals to the corresponding dialkyl ethers [Howard & Brown *J Org Chem* **26** 1026 1961], hydrogenation of heterocyclic compounds e.g. 1,2,3,4-tetrahydroisoquinoline-3-carboxylic acid to *cis*-decahydroisoquinoline-3-carboxylic acid [40psi,  $\sim 25^\circ$ , 3 hours;

Rapala et al. *J Am Chem Soc* **79** 3770 1957], nicotinic acid to hexahydronicotinic (nipecotinic) acid without decarboxylation [2 atmospheres, ~25°, 4 hours; Freifelder *J Org Chem* **27** 4046 1962, Freifelder *J Org Chem* **28** 602, 1135 1963], complete hydrogenation of quinoxaline [Broadbent et al. *J Am Chem Soc* **82** 189 1960], reduction of the 4,5-double bond of pyrimidine nucleosides and nucleotides [Cohn & Doherty *J Am Chem Soc* **78** 2863 1956], dehydrogenation e.g. hydrogen transfer from hexahydrohexahelicene to benzene in order to provide hexahelicene and cyclohexane in 73% yield [Newman & Lednicer *J Am Chem Soc* **78** 4765 1956, cf also Anderson & Anderson *J Org Chem* **22** 1197 1957], and hydrogenation of anilines to cyclohexylamine with little hydrogenolysis [Freifelder et al. *J Org Chem* **30** 2485 1965].

**Rhodium on carbon (5% and 10% Rh loading).** A 10%-Rh, 0.1%Pd on carbon catalyst has been prepared from a mixture of  $\text{RhCl}_3 \cdot 3\text{H}_2\text{O}$  (5.26g),  $\text{PdCl}_2$  (0.34g) and carbon (18g, e.g. Darco G-60, or any acid washed charcoal can be used) in  $\text{H}_2\text{O}$  (200ml) is rapidly stirred and heated to 80°. Lithium hydroxide hydrate (2.7g) dissolved in  $\text{H}_2\text{O}$  (10ml) is added in one lot and heating is stopped but the mixture is stirred overnight, filtered, and washed with 0.5% v/v aqueous AcOH (100ml) then dried *in vacuo* at 65° to give 20.6-21g of catalyst. One gram of this catalyst will absorb 0.0022-0.0028 mole of  $\text{H}_2$  in aqueous suspension. It has been used, for example, in the reduction of *p*-aminobenzoic acid to a mixture of *cis*- and *trans*- 4-aminocyclohexanecarboxylic acid [at 50psi/~25°/~24 hours] which was converted to 2-azabicyclo[2.2.2]octan-3-one (3-isoquinuclidone) in 81-84% yield. *The point was made that when fresh catalyst is added to the reaction vessel, it should be moistened with the solvent; and that hydrogen should be removed by evacuation at the end of the reduction in order to avoid having a hydrogen-air atmosphere which could be explosive.* [Pearlman *Org Synth Coll Vol V* 670 1973, Pearlman *Tetrahedron Lett* 1663 1967.] It is a useful catalyst (even without the trace of Pd) for the reduction of nitriles to primary amines, and the hydrogenation of aromatic and pyridine rings, aldehydes and sugars. It is superior to  $\text{Rh-Al}_2\text{O}_3$  for reduction of the pyridine ring and larger amounts of catalyst have to be added to overcome the poisoning effect of the piperidine products [Freifelder et al. *J Org Chem* **27** 284 1962]. It is useful for the stepwise reduction of 3-*H*-pyrrolizine to 3,4-dihydropyrrolizine, and then to the fully reduced pyrrolizidine [Schweizer & Light *J Am Chem Soc* **86** 2963 1964]; and in boiling norbornadiene complete conversion to a mixture of dimers and trimers occurs [Mrowca & Katz *J Am Chem Soc* **88** 4012 1966]. A comparison of the rates of hydrogenation with 5% Pd/C, 5%Pt/C, 5% Rh/C and 5% Ru/C catalysts towards reduction of C=C, C=O, C-OH and benzene rings in various solvents at 25° and initial  $\text{H}_2$  pressure of 1 atmosphere were made by Breitner, Roginski & Rylander [*J Org Chem* **24** 1855 1959].

**Ruthenium (1,5-Cyclooctadiene)(1,3,5-cyclooctatetraene) [Ru(cod)(cot)]** [127382-91-6] **M 315.4, m 88-94°, 92-94°.** For preparation see Chapter 5, Metal-Organic Compounds. Purify the ruthenium complex (~ 0.3g) by dissolving it in *n*-pentane (~ 70ml) and filtering the solution through a column of alumina (Merck, Brockman activity II—III, 20cm). Collect the yellow band and reduce its volume to ~ 5ml then cool it at -70°, preferably under  $\text{N}_2$  or Ar. After ~ 6 hours collect the yellow solid (~ 0.2g) and dry it *in vacuo*. Recrystallisation from *n*-pentane results in a 75-85% recovery. [Itoh et al. *J Organomet Chem* **272** 179 1984, Petrici et al. *J Chem Soc, Dalton Trans* 1961 1980]. It is a highly selective catalyst for amine alkylations in which, unlike other ruthenium catalysts, alkylates the amino group of e.g., 2-, 3-, or 4- aminopyridines as well as 2-aminopyrimidine to give the respective monoethylamino derivatives by using EtOH in high yields with very little, if any, of diethylamino derivatives. Thus it stops almost exclusively at monoalkylation. [Watanabe et al. *J Org Chem* **61** 4214 1996]. It also catalyses the dimerisation of 2,5-norbornadiene *via* [2+2] cycloaddition reactions and eyne generation in the presence of dimethylfumarate in THF at 40° for 1 hour to form the “cup-shaped” molecule pentacyclo[6.6.0.0<sup>2,6</sup>.0<sup>3,13</sup>.0<sup>10,14</sup>]tetradeca-4,11-diene in 96% yield. An alicyclic compound which would otherwise require many steps to synthesis. [Mitsudo et al. *J Am Chem Soc* **121** 1839 1999.]

**Ruthenium metal (Ru) used in preparing the following catalysts has** [7440-18-8] **M 101.1, m 2310°, b 3900°, d<sup>20</sup> 12.5].**

**Ruthenium-aluminium oxide/hydroxide catalyst (~2.5wt% Ru loading).** The preparation of this catalyst is described in Chapter 8 because it consists of Ru(0) nanoparticles encapsulated in an aluminium oxy-hydroxide

matrix and is superior as a catalyst to the commercially available material. It is a recyclable catalyst for the efficient oxidant-free alcohol dehydrogenation (compare with  $\text{RuCl}_3 \cdot 3\text{H}_2\text{O}$  or  $x\text{H}_2\text{O}$  above) of a wide range of alcohols including benzylic alcohols, secondary long chain fatty alcohols, secondary allylic alcohols, 3-pyridylmethanol, 2-thiophenylmethanol and 4-methylthiobenzyl alcohol in dry toluene to the respective carbonyl compounds under argon at dehydrogenation temperatures of  $80^\circ$  to  $110^\circ$ , using 3-6 mol% of Ru, in yields of  $>90\%$ . It is 5 times superior to the commercial  $\text{Ru}/\text{Al}_2\text{O}_3$  catalyst and is re-usable, after washing with  $\text{Me}_2\text{CO}$  and drying in air, at least 10 times without loss of activity. [Kim, Park and Park *Org Lett* **8** 2543 2006.]

**Ruthenium on alumina (5% Ru reduced loading).**  $\text{Ru}/\text{Al}_2\text{O}_3$  can be freshly prepared by treating alumina powder with a ruthenium compound, e.g.  $\text{RuCl}_3 \cdot x\text{H}_2\text{O}$  in  $\text{H}_2\text{O}$ ,  $\text{Ru}(\text{acac})_3$  or  $\text{Ru}_3(\text{CO})_{12}$  in \*benzene (to give a solid that is  $\sim 5\text{wt}\%$  in Ru), followed by evaporating the solvent and reducing the solid in a stream of  $\text{H}_2$  at  $500^\circ$ . This catalyst will oxidise dibenzothiophene or dibenzothiophene oxide (Ru molar ratio 11:1) in \* $\text{C}_6\text{H}_6$  to dibenzothiophene sulfone at  $100^\circ$  in air (70 atmospheres) for 12 hours in 97% yield. Slow poisoning of Ru occurs during the oxidation. The activity of the catalyst depends on the preparation, but is much more active than the following catalysts in the sequence  $\text{Ru}_3(\text{CO})_{12} > \text{Ru}(\text{acac})_3 > \text{RuCl}_3 \cdot x\text{H}_2\text{O}$ .  $\text{Ru}/\text{C}$  or  $\text{Ru}/\text{SiO}_2$ , and the commercially available  $\text{Ru}/\text{Al}_2\text{O}_3$  are less active than the above preparation.  $\text{Pt}/\text{Al}_2\text{O}_3$  or  $\text{Ir}/\text{Al}_2\text{O}_3$  gives only a trace of the sulfone, and  $\text{Pd}/\text{Al}_2\text{O}_3$ ,  $\text{Rh}/\text{Al}_2\text{O}_3$ ,  $\text{Os}/\text{Al}_2\text{O}_3$  or  $\text{Cu}/\text{Al}_2\text{O}_3$  are totally inactive. It should be noted that with this catalyst oxidation of  $n\text{-Bu}_2\text{S}$  (0.2M in \* $\text{C}_6\text{H}_6$ ; in the ratio of Ru/sulfide 21:1) provides 44 and 9% molar yields of  $n\text{-Bu}_2\text{SO}$  and  $n\text{-Bu}_2\text{SO}_2$  respectively, with full recovery of unreacted  $n\text{-Bu}_2\text{S}$ . [Ledlie & Howell *Tetrahedron Lett* 785 1976.]

**Ruthenium on carbon (5% Ru loading).** A comparison of the rates of hydrogenation with 5%  $\text{Pd}/\text{C}$ , 5%  $\text{Pt}/\text{C}$ , 5%  $\text{Rh}/\text{C}$  and 5%  $\text{Ru}/\text{C}$  catalysts towards reduction of  $\text{C}=\text{C}$ ,  $\text{C}=\text{O}$ ,  $\text{C}-\text{OH}$  and benzene rings in various solvents at  $25^\circ$  and initial  $\text{H}_2$  pressure of 1 atmosphere were made by Breitner, Roginski & Rylander [*J Org Chem* **24** 1855 1959.]  $\text{Ru}/\text{C}$  is a superior catalyst than  $\text{Pd}/\text{C}$  and  $\text{Pt}/\text{C}$  for the hydrogenation of ketones in basic and neutral media and in some cases (e.g. tetramethylcyclobutan-1,3-dione at 1000-1500psi of  $\text{H}_2/125^\circ/1$  hour)  $>90\%$  yields of diols were obtained [Hseck et al. *J Org Chem* **27** 700 1961],  $\text{C}=\text{C}$  (with some selectivity), and acetylenes were reduced to alkanes [Berkowitz & Rylander *J Org Chem* **24** 708 1959], nitrobenzene was reduced to hydrazobenzene in alcoholic  $\text{KOH}$  in 80% yield without over-reduction to aniline [Peitra & Res *Ann Chim (Rome)* **48** 299 1958], and 10%  $\text{Ru}/\text{C}$  was better than other catalysts for reducing phenolic rings in polycyclic compounds [Walton et al. *J Am Chem Soc* **78** 4760 1956].

**Ruthenium(IV) oxide ( $\text{RuO}_2$ )** [12036-10-1] **M 133.1,  $d^{25}$  6.97.** The dioxide is prepared by heating Ru metal in a silica boat in a silica combustion tube placed horizontally in a furnace, and a slow stream of dry oxygen is allowed to flow over it while it is being heated at  $1000^\circ$  for 24 hours. During heating a little of the oxide volatilises and small crystals may condense on the walls at the downstream end of the tube. A 90-95% yield of polycrystalline  $\text{RuO}_2$  remains in the boat. It forms tetragonal dark blue crystals which are insoluble in organic solvents,  $\text{H}_2\text{O}$  and acids but soluble in fused alkali. [Schäfer & Heitland *Z Anorg Allgem Chem* **304** 249 1960, Schäfer et al. *Z Anorg Allgem Chem* **319** 327 1963, Rogers et al. *Inorg Synth* **XIII** 137 1972, **XXX** 96 1995.] It is a good catalyst, no doubt by being converted to the metal *in situ* which promotes catalysis. It has been used for the reduction of aromatic rings to cyclohexane rings [e.g. at 1350psi/147°/16 hours, or 70psi/90°/1 hour; Freifelder & Stone *J Am Chem Soc* **80** 5270 1958, Freifelder & Stone *J Org Chem* **26** 3805 1961, Freifelder & Stone *J Org Chem* **27** 3568 1962, Johnson et al. *J Am Chem Soc* **84** 2181 1962], and phenolic compounds to the respective cyclohexanols [Rapala & Farkas *J Org Chem* **23** 1404 1958, Counsel *Tetrahedron* **15** 202 1961, Ireland & Schiess *J Org Chem* **28** 6 1963]; but with some enones only the double bond was reduced without affecting the carbonyl function [Rapala & Farkas *J Am Chem Soc* **80** 1008 1958].

## HOMOGENEOUS METAL CATALYSTS

**Allylpalladium(II) chloride dimer {di- $\mu$ -chloro- $\eta^3$ -allylpalladium(II),  $[\text{Pd}(\eta^3\text{-C}_3\text{H}_5)\text{Cl}]_2$ }** [12012-95-2] **M 365.9, m 155-156°(dec.),  $\sim 160^\circ$ (dec.).** It is prepared from allyl chloride,  $\text{PdCl}_2$ , and  $\text{NaCl}$  in the presence of carbon monoxide and is extracted into  $\text{CHCl}_3$ , washed with  $\text{H}_2\text{O}$ , dried ( $\text{CaCl}_2$ ), filtered and evaporated to

dryness to give the complex as yellow crystals (use a very efficient fume cupboard). These can be used as such, but by recrystallisation from a mixture of  $\text{CH}_2\text{Cl}_2$  and hexane analytically pure catalyst can be obtained (**m 155-156° dec.**). It also crystallises from benzene and is soluble in MeOH,  $\text{Et}_2\text{O}$ ,  $\text{Me}_2\text{CO}$  and  $\text{CHCl}_3$ . The  $^1\text{H}$  NMR ( $\text{CDCl}_3$ ) has two doublets at  $\delta_{\text{H}}$  3.03 (anti  $\text{CH}_2$ ,  $J = 12.0\text{Hz}$ ) and 4.10 (syn  $\text{CH}_2$ ,  $J = 7.1\text{Hz}$ ), and a triplet at  $\delta_{\text{H}}$  5.48 (CH) 2:2:1. [Tatsuno et al. *Inorg Synth* **28** 342 1990, Shaw *Proc Chem Soc* 247 1960, Hüttel et al. *Chem Ber* **94** 766 1961, Dent et al. *J Chem Soc* 1585 1964.]

It is a labile Pd complex, but is very useful for preparing a variety of stable catalytic Pd(0) complexes with a wide selection of ligands to perform stereoselective reactions. A particularly useful ligand is *cis,cis,cis*-1,2,3,4-tetrakis(diphenylphosphinomethyl)cyclopentane [Tedicyp] [Laurenti et al. *J Org Chem* **66** 1633 2001]. The catalyst with this ligand is typically prepared under argon, in a Schlenk type system, by dissolving  $[\text{Pd}(\eta^3\text{-C}_3\text{H}_5)\text{Cl}]_2$  (4.2mg, 11.6 $\mu\text{mol}$ ) and Tedicyp (20mg, 23.2 $\mu\text{mol}$ , see [333380-86-2] in Part 2) in dry tetrahydrofuran (10ml) and stirring for 10 minutes before use. It gives a solution of 2.32 $\mu\text{mol}/\text{ml}$  of catalyst, which has  $^{31}\text{P}$  NMR (162MHz,  $\text{CDCl}_3$ ) with  $\delta$  at 25 (vbrs, width 80Hz), 19.4 (vbrs, width 110Hz). It is very efficient in catalysing allylic substitutions with high turnover numbers (up to 9,800,000) and frequencies (up to 190,000  $\text{h}^{-1}$ ).

$[\text{Pd}(\eta^3\text{-C}_3\text{H}_5)\text{Cl}]_2$  is used for cross-coupling reactions [Wallow & Novak *J Org Chem* **59** 5034 1994], the Heck reactions with olefins [Fall et al. *Synthesis* 1683 2007] and enol ethers [Battace et al. *Eur J Org Chem* 3122 2007], and the formation of ethyl cyclopropane carboxylates with ethyl diazoacetate and olefins [Armstrong *J Org Chem* **31** 618 1966] to name a few.

**Bis(acetonitrile)dichloropalladium(II)  $[\text{Pd(II)Cl}_2(\text{MeCN})_2$  complex]** [14592-56-4] **M 259.4, dec > 200°.** It is prepared by mixing anhydrous  $\text{PdCl}_2$  (1mol) and MeCN (2mols) until clear and precipitating the complex with petroleum ether. *Alternatively*, MeCN (3ml) is added to a solution of  $\text{K}_2\text{PdCl}_4$  (2g) in  $\text{H}_2\text{O}$  (40ml) and stirred at  $\sim 25^\circ$  until separation of yellow prisms is complete (could take up to 8 days). Filter these off, wash them with  $\text{H}_2\text{O}$ , EtOH,  $\text{Et}_2\text{O}$ , and dry them *in vacuo*. Further purification, if necessary, can be achieved by dissolving it in anhydrous solvents (which are not likely to displace the ligands) such as MeOH, toluene, THF, precipitating it with light petroleum, and drying the solid *in vacuo*. It is also soluble in  $^*\text{C}_6\text{H}_6$  to the extent of 100mg/4.0ml. In addition to its use as a source of metal for palladium complexes and catalysts such as  $\text{PdCl}_2(\text{R})\text{-BINAP}$  [115826-95-4] (see Ozawa et al. *Organometallics* **12** 4188 1993), it also catalyses Overman rearrangements (aza-Claisen rearrangement, Overman & Garpenher *Org React* **66** 1 2005, see below). A particularly interesting example of the rearrangement being the high diastereoselectivity which is directed by a  $\delta$ -methoxymethylether group in  $\delta$ -substituted allylic trichloroacetamides, where the ether oxygen atoms coordinate with Pd to direct the orientation of the reaction [Jamieson & Sutherland *Tetrahedron* **63** 2123 2007]. The IR spectrum shows, from the RCN and M-X vibrations, that the MeCN groups have a *trans*-planar configuration, unlike related Pt complexes which have the *cis*-configuration. The IR (Nujol) has  $\nu_{\text{max}}$  at 1417m, 1358m,  $\sim 1160\text{wbr}$ , 1030m, 970wbr  $\text{cm}^{-1}$  [Walton *Spectrochim Acta* **21** 1795 1965].

**Bis(benzonitrile)dichloropalladium(II)  $[\text{Pd(II)Cl}_2(\text{PhCN})_2$  complex]** [14220-64-5] **M 383.6, m 131°.** It is prepared and purified in the same manner as the  $\text{Pd(II)Cl}_2(\text{MeCN})_2$  complex above. *Alternatively*, anhydrous  $\text{PdCl}_2$  (2g) suspended in benzonitrile (50ml) is stirred at  $100^\circ$ . Most of the  $\text{PdCl}_2$  dissolves after 20 minutes to give a red solution which is filtered while hot, and the hot filtrate is poured slowly into petroleum ether (b 40-60°). The light yellow complex which crystallises out is collected, washed with petroleum ether and dried *in vacuo* (4.0g, 93%). It can be purified further by dissolving it in the minimum volume of  $^*\text{C}_6\text{H}_6$ , and the hot red solution is filtered and diluted with petroleum ether to allow the complex to crystallise out. [Doyle, Slade and Jonassen *Inorg Synth* **VI** 218 1960, Kharasch et al. *J Am Chem Soc* **60** 882 1938.]  $\text{Pd(II)Cl}_2(\text{PhCN})_2$  is an effective catalyst for the regiospecific cycloaddition cross-coupling reactions of aziridines with carbodiimides to form imidazolidin-2-imines in 40-94% yields [Baeg & Aler *J Org Chem* **57** 157 1992]. In the presence of a Buchwald ligand (e.g. 2-di-*t*-butyl-2'4'6'-tri-*i*-propylbiphenyl) in  $\text{CH}_2\text{Cl}_2$ , it catalyses the stereoselective formation of  $\alpha$ -*O*-glycosides from protected glycalimides with a variety of phenols or protected glycosides in high yields at room temperature [Schuff, Mercer and Nguyen *Org Lett* **9** 3174 2007]. Also, in the presence of Xantphos {4,5-bis(diphenylphospheno)-9,9-dimethylxanthene, see [161265-03-8]}, it promotes Stille-type cross coupling reactions between alkynyl- $\text{SnBu}_3$  and  $\alpha$ -halo esters or amides under neutral conditions in THF to provide 3-alkynoates or alkynoamides in high yields without homocoupling of the alkynyltin reagents [Shi et al.

*Chem Commun* 2342 2007]. The IR spectrum shows, from the RCN and M-X vibrations, that the PhCN groups have a *trans*-planar configuration, unlike related Pt complexes which have the *cis*-configuration. The IR (Nujol) has  $\nu_{\max}$  at 3096w, 3057w, 3033w, 3002vwm 2287s, 1480m-w, 1440s, 760s, 678s  $\text{cm}^{-1}$  [Walton *Spectrochim Acta* **21** 1795 1965].

**Bis(cyclopentadienyl)cobalt(II) (cobaltocene,  $\text{C}_5\text{H}_5\text{CoC}_5\text{H}_5$ ,  $\text{Cp}_2\text{Co}$ )** [1277-43-6] **M 189.1, m 173-174° (under  $\text{N}_2$ ), 176-180°(dec)**.  $\text{Cp}_2\text{Co}$  is prepared by adding anhydrous  $\text{CoCl}_2$  (0.5 mol, prepared by dehydrating the salt at 200° in a vacuum) to a cold solution of 1 mol of sodium cyclopentadienide [4984-82-1] in THF (see below), stirring for 2 hours, removing the solvent and subliming the product directly from the mixture at 150-200°/10<sup>-4</sup> mm in 75-80% yields; or by using  $\text{CoBr}_2$  and refluxing for 3 hours. *Alternatively* (method of Wi et al. *J Inorg Nucl Chem* **2** 110 1956), cyclopentadiene (42ml, 0.5mol) and  $\text{Et}_2\text{NH}$  (~100ml, 1mol) are added slowly to  $\text{CoBr}_2$  (0.25mol) with cooling, and then stirring at ~25° for 6-8 hours; the solvent is removed *in vacuo* and the residue is extracted with petroleum ether, filtered, and the solvent is evaporated to give dark violet monoclinic crystals m 171-173°.

When anhydrous *cobalticenium chloride* (5g, prepared from sodiocyclopentadienide and anhydrous  $\text{CoCl}_2$  in THF as above but stirring at the boiling point for 2 hours, evaporating the solvent and adding 1:1 hydrochloric acid) in dry 1,2-dimethoxyethane (150ml) is treated with excess of  $\text{NaBH}_4$  (5g, or LAH) in small portions while stirring ( $\text{H}_2$  is evolved), then warm to 60°, or add  $\text{H}_2\text{O}$  (2ml), the mixture becomes deep wine-red in colour. After 1 hour the solution is filtered, the solid is washed with  $\text{Et}_2\text{O}$  and the filtrate is evaporated to a few mls *in vacuo*, diluted with petroleum ether (b ~40-60°), concentrated again and applied on to an alumina column (20 x 2cm). The first red band that elutes with petroleum ether is evaporated, and the residue is sublimed *in vacuo* onto a cold finger at -70° to give wine-red crystals of  $\pi$ -cyclopentadienyl(cyclopentadiene)cobalt [ $\pi$ - $\text{C}_5\text{H}_5\text{CoC}_5\text{H}_6$ ] (yield 80%) with **m 98-99°**, has **M 193** (ebulioscopic in  $\text{C}_6\text{H}_6$ ), and a UV with  $\lambda_{\max}$  nm( $\epsilon$ ) at 263.8 (1.78 x 10<sup>5</sup>), 327 (9.27 x 10<sup>3</sup>), and 395 (5.27 x 10<sup>3</sup>). It has a camphorous odour and should be kept under argon as it is stable in air for only a few minutes, but for a few hours in degassed solutions of organic solvents. [Green et al. *J Chem Soc* 3753 1959.]

**Sodium cyclopentadienide** [4984-82-1] is prepared under  $\text{N}_2$  by slowly adding cyclopentadiene (37ml) to a cooled suspension of sodium (10g) in THF (150ml) when  $\text{H}_2$  evolves rapidly and a solution of sodium cyclopentadienide is formed. The colour of the solution depends on the purity of the solvent and the amount of air in the system. Clear pale orange or red solutions are obtained with good conditions, but dark red to purple colours are obtained in the presence of traces of air without, however, seriously decreasing the yields of the complexes formed. [Wilkinson et al. *J Inorg Nucl Chem* **2** 95 1956, Reynolds & Wilkinson *J Nuclear Chem* **9** 86 1959, *Beilstein* **16** IV 1698.] It is available commercially as a 2.0M solution in THF as sodium cyclopentadienylide. The dry solid is **pyrophoric**.

$\text{Cp}_2\text{Co}$  catalyses the condensation of acetylene and mono-substituted acetylenes (but not di-substituted acetylenes) with nitriles at 150° to give 30-70% yields of substituted pyridines [Wakatsuki & Yamazaki *Synthesis* **26** 1976]. This synthesis has been applied successfully in a 2+2+2 cycloaddition reaction between di(trimethylsilyl)propargyl ether and acetonitrile to prepare a pyridine intermediate which was used to obtain pyridoxine (vitamin B<sub>6</sub>) [Greiger et al. *Helv Chim Acta* **67** 1274 1984].

*The above are general procedures for preparing* **Bis(cyclopentadienyl)chromium (II) [ $\text{Cp}_2\text{Cr}$ ]** [1271-24-5] **M 182.2, m 168-170°** [*Beilstein* **16** IV 1774, Wilkinson et al. *J Inorg Nucl Chem* **2** 95 1956.], **Bis(cyclopentadienyl)magnesium(II) [ $\text{Cp}_2\text{Mg}$ ]** [1284-72-6] **M 154.5, m 176-178°, 180°(dec)** [*Beilstein* **16** IV 1695, Wilkinson et al. *J Inorg Nucl Chem* **2** 95 1956, Barber *J Inorg Nucl Chem* **4** 373 1957], and **Bis(cyclopentadienyl)vanadium(II) [ $\text{Cp}_2\text{V}$ ]** [1277-47-0] **M 181.1, m 165-167°** [*Beilstein* **16** IV 1771]. [Wilkinson et al. *J Inorg Nucl Chem* **2** 95 1956.]

**Bis(dibenzylideneacetone)palladium(0) [ $\text{Pd}(\text{dba})_2$ ]** [32005-36-0] **M 575.0**. The precise nature of this complex is not clear but a product with reproducible stoichiometry [ $\text{C}_{34}\text{H}_{28}\text{O}_2\text{Pd}$ ] is prepared first by stirring  $\text{PdCl}_2$  (8.87g, 50mmol) and  $\text{NaCl}$  (2.92g, 50mmol) in MeOH (250ml) at 25° for 16 hours, filtering through a cotton wool plug, diluting the filtrate to ~1.5L with MeOH, and heating to 60° to form  $\text{Na}_2[\text{PdCl}_6]$ . Dibenzylideneacetone (36.5g, 150mmol, see [538-58-9], cf Organic Compounds in Chapter 4) is then added to this warm solution stirred solution of  $\text{Na}_2[\text{PdCl}_6]$ , and stirring for a further 15 minutes and anhydrous  $\text{NaOAc}$  (75g) is added. The reaction begins immediately, heating is stopped but stirring is continued until the reaction

cools to room temperature. The dark brown precipitate that is formed is washed with MeOH (2 x 25ml), H<sub>2</sub>O (5 x 50ml), Me<sub>2</sub>CO (5 x 15ml) and dried in air to constant weight (923g, 80% yield) [Rettig & Maitlis *Inorg Synth* **17** 134 1977]. See the related complex **Pd(dm-dba)<sub>2</sub>** below for the structure of these complexes in CDCl<sub>3</sub> solution. Catalysis with these complexes occurs at room temperature in the presence the ligand DIPHOS [dpe, 1,2-bis-[diphenylphospheno]ethane (see [1663-45-2] in Metal-Organic Compounds in Chapter 5) which is more efficient than with the less stable Pd(PPh<sub>3</sub>)<sub>4</sub>. It is used for the synthesis of allylic substituted cyclopentadienes [Fiaud & Mallerson *Tetrahedron Lett* **21** 4437 1980] and is a homogeneous catalyst [Fiaud & Mallerson *Tetrahedron Lett* **22** 1399 1981, Black *Aldrichimica Acta* **15** 13 1982] which catalyses the alkylation of a variety of nucleophiles under mild conditions [Ferroud et al. *Tetrahedron Lett* **25** 4379 1984.] Together with cyclic thiourea ligands such as *1N,3N*-bis(1,4-di-*tert*-butylphen-2-yl)imidazoline-2-thione, it catalyses efficient aerobic oxidations of alcohols to aldehydes and ketones in toluene in yields >95% [Yang et al. *Synlett* 3057 2006.]

**Bis(3,5,3',5'-dimethoxybenzylideneacetone)palladium(0)** {**bis[1,3-bis(3',5'-methoxybenzylidene)acetone palladium(0), [Pd(dm-dba)<sub>2</sub>]**} [811862-77-8] **M 815.2, m 170-178°**. This Pd catalyst with the dimethoxy-dba ligand is prepared and purified as for Pd(dba)<sub>2</sub> above but using 1,5-bis(3',5'-dimethoxyphenyl)penta-1*E*,4*E*-diene-3-one (dm-dba, [39777-58-7]),\* instead of dba. *Alternatively*, Fairland's method can be used, whereby NaCl (3.3g, 56mmol) is added to PdCl<sub>2</sub> (5g, 28mmol) in MeOH (140ml) and stirred under argon at 25° for 24 hours, filtered and evaporated *in vacuo* to half its volume. To this solution at 60° is added dm-dba (20.6g, 88mmol), stirred at 60° for 15 minutes, anhydrous NaOAc (42.3g) is added and allowed to cool to 25°. The mixture is stirred at 25° for 2 hours as the dark red precipitate of the complex separates. The solid is collected, washed with MeOH (2 x 100ml), H<sub>2</sub>O (2 x 100ml) then Me<sub>2</sub>CO (2 x 20ml), partially dried by suction, placed in a Schlenk flask and stirred under a flow of N<sub>2</sub> overnight to give the complex as a maroon/purple microcrystalline solid (11g 68%). It was found to have higher activity than those of the complex with the unsubstituted ligand [i.e. Pd(dba)<sub>2</sub>] in Suzuki-Miyaura cross-coupling reactions of organic halides with arylboronic acids. *Note* that a solution of the complex, as in the case of Pd(dba)<sub>2</sub> above, in CDCl<sub>3</sub>, was found to contain one dm-dba ligand coordinated to Pd and one free in solution. Addition of two equivalents of Ph<sub>3</sub>P to the complex in THF and (CD<sub>3</sub>)<sub>2</sub>CO resulted in the formation of the  $\eta^2$ -(dm-dba)Pd<sup>(0)</sup>(Ph<sub>3</sub>P)<sub>2</sub> complex as shown by the presence of two phosphorous doublets in the <sup>31</sup>P NMR spectrum. [Fairlamb et al. *Org Lett* **6** 4435 2004, cf *Handbook of Organopalladium for Organic Synthesis* Negishi ed. Wiley, Hoboken NJ 2002, ISBN 0-471-31506-0.] For **Pd<sub>2</sub>(p,p'-dimethoxy-dba)<sub>3</sub>** see **Pd<sub>2</sub>(dba)<sub>3</sub>**(CDCl<sub>3</sub>) below.

**Bis(di-*tert*-butyl(4-dimethylaminophenyl)phosphine)dichloropalladium(II)** [(A-<sup>ta</sup>Phos)<sub>2</sub>PdCl<sub>2</sub>] [887919-35-9] **M 708.1**. The catalyst is prepared by mixing one equivalent of di-*tert*-butylphosphine, 1 equivalent of 4-dimethylaminophenyl bromide and 1 mol% of Pd<sub>2</sub>dba<sub>3</sub> at 90° for 12 hours. The mixture is cooled to 25°, filtered through a short silica gel plug, the plug is rinsed with toluene and the combined filtrate is evaporated under vacuum. The residue is dissolved in tetrahydrofuran and 0.35 equivalents of solid PdCl<sub>2</sub>(COD) are added followed by stirring at 25° for 12 hours. The product is filtered off, and the yellow solid collected is washed with pentane and dried *in vacuo* to give ~85% yield of (A-<sup>ta</sup>Phos)<sub>2</sub>PdCl<sub>2</sub> which provides the correct microanalysis for C and H. It has <sup>1</sup>H NMR (CD<sub>2</sub>Cl<sub>2</sub>) with  $\delta$  at 7.5 (s br, 4 H, ArH), 6.42 (d, *J* = 8.3Hz, 4H, ArH), 2.75 (s, 12H, NMe<sub>2</sub>), 1.38 (t, *J*<sub>PH</sub> = 6.6Hz, 36H, *tert*-Bu); and <sup>31</sup>P NMR (CDCl<sub>3</sub>) with  $\delta$  at 52.1. Similarly **bis[di-*tert*-butyl(*p*-trifluoromethylphenyl)phosphine)dichloropalladium(II)** can be prepared which has <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 8.0 (s br, 4 H, ArH), 7.59 (d, *J* = 7.9Hz, 4H, ArH), 1.62 (t, *J*<sub>PH</sub> = 7.0Hz, 36H, *tert*-Bu); <sup>31</sup>P NMR (CDCl<sub>3</sub>) with  $\delta$  at 55.7; and <sup>19</sup>F NMR (CDCl<sub>3</sub>) with  $\delta$  at -63.4. Also **bis[di-*tert*-butyl(phenyl)phosphine)dichloro-palladium (II)** was prepared in the same way and has <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 7.9-7.8 (m, 4 H, ArH), 7.36-7.30 (m, 6H, ArH), 1.56 (t, *J*<sub>PH</sub> = 6.9Hz, 36H, *tert*-Bu); <sup>31</sup>P NMR (CDCl<sub>3</sub>) with  $\delta$  at 54.4. These are air-stable active catalysts for Suzuki-Miyaura cross-coupling with aryl halides including 5- and 6-membered heteroaryl halides in the presence of bases such as K<sub>2</sub>CO<sub>3</sub>, Na<sub>2</sub>CO<sub>3</sub>, KOAc, K<sub>3</sub>PO<sub>4</sub> with high turn over numbers (TON) of 100 to 100,000 [Guram et al. *Org Lett* **8** 1787 2006].

**2-[Bis(2,4-di-*tert*-butylphenoxy)phosphinoxy]-3,5-di(*tert*-butyl)phenyl-palladium(II) chloride dimer {Bedford's Catalyst, 2[bis(2,4-di-*tert*-butylphenoxy)phosphino-κP-oxy]-3,5-di-*tert*-butylphenyl-κC]di-μ-chloro-dipalladium}** [217189-40-7] **M 1575.6**. The thoroughly characterised Bedford catalyst, an *ortho*-metallated dimer, is prepared in 96% yield from the bulky tris(2,4-di-*tert*-butylphenyl)phosphite and PdCl<sub>2</sub>. The catalyst is remarkably stable to moisture and air, showing no signs of deterioration after several weeks, can be kept in air for at least 6 months and is not decomposed on heating at 130° in toluene for 24 hours. It is very efficient for coupling aryl halides with either arylboronic acids (Suzuki reaction) or aryltin compounds (Stille reaction) in solvents such as toluene or tetrahydrofuran to form biaryls with very high turn over numbers (TONs up to 1,000,000 moles of product/mol of Pd) at even very low Pd catalyst concentrations, and in moderate to high yields [Albisson et al. *J Chem Soc, Chem Commun* 2095 1998]. Bedford's catalyst also proved to be extremely active in the arylation (with e.g. aryl bromides) of alkenes (e.g. acrylate, styrene) with turn over numbers of up to 5,750,000 (mole product/mol Pd), and turn over frequencies of 300,000 (mol product/mol Pd/hour) [Albisson et al. *Tetrahedron Lett* **39** 9793 1998].

***R,R*-(+) and *S,S*-(-) *N,N'*-Bis(3,5-di-*tert*-butylsalicylidene)-1,2-*trans*-cyclohexanediamine Manganese(III) Chloride [Jacobsen's catalyst] [*R,R*-138124-32-0, *S,S*-135620-04-1]** **M 635.2, m 330-332°, 324-326°, [α]<sub>D</sub><sup>20</sup> ±580° (c 0.01, EtOH)**. This complex is commonly referred to as the Mn(salen) catalyst. If this enantioselective catalyst for the epoxidation of olefins requires purification, the brown powder (*ca* 6g) is dissolved in CH<sub>2</sub>Cl<sub>2</sub> (100ml), extracted with H<sub>2</sub>O (2 x 100ml), the organic phase is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated to dryness and the solid residue is dried in a vacuum desiccator over CaSO<sub>4</sub>. [Jacobsen et al. *J Am Chem Soc* **113** 7063 1991]. For large scale preparation see Larrow et al. *J Org Chem* **59** 1939 1994, Deng & Jacobsen *J Org Chem* **57** 4320 1992. For epoxidation of indenenes see Hughes et al. *J Org Chem* **62** 2222 1997, and for the epoxidation of isoflavones see Adams et al. *Tetrahedron: Asymmetry* **9** 1121 1998, Chang et al. *Tetrahedron* **58** 6939 1993, Jacobsen et al. *Tetrahedron* **50** 4323 1994, Jacobsen's Asymmetric Catalytic Epoxidation of Unfunctionalized Olefins in *Catalytic Asymmetric Synthesis Vol I*, p. 159, Ojima Ed. VCH, NY 1993 (see also Bibliography in Part 2), and for enantioselective hydrolysis of epoxides see Furrow et al. *J Org Chem* **63** 6776 1998.

On a large scale, the brown *R,R*-(+)- catalyst (300-400g) is washed with hot H<sub>2</sub>O (2 x 800ml, at 50°), and dried on the filter for 2 hours then under vacuum (50-100mm) at 60-70° for 12 hours, m 324-326° (product of acceptable purity has m ~320°), [α]<sub>D</sub><sup>23</sup> +580° (c 0.01, EtOH), and IR with ν<sub>max</sub> (KBr) at 1535, 1612, 2950-2958 cm<sup>-1</sup>. [Larrow et al. *J Org Chem* **59** 1939 1994.]

The related chiral complexes of Al(II)-Cl [*R,R* 250611-13-3; *S,S* 307926-51-8], of Cr(III)-Cl [*R,R* 164931-83-3; *S,S* 219143-92-7] and of the Co(III)-Cl [*R,R*-(+) 176762-62-5; *S,S*-(-) 188264-84-8] (Tokunaga et al. *Science* **277** 936 1977, Jacobsen et al. *Tetrahedron Lett* **38** 733 1997) are similarly prepared and are available commercially.

NOTE: It was shown that addition of pyridine *N*-oxides, and particularly 4-(3-phenylpropyl)pyridine-1-oxide (P<sub>3</sub>NO, see [34122-28-6] in Part 2), stabilise and improve the catalytic activity of Mn(salen) catalysts [Larrow & Jacobsen *J Am Chem Soc* **116** 12129 1994, Srinivan et al. *J Am Chem Soc* **108** 2309 1986, Senanayake *Aldrichimica Acta* **31** 3 1998.] Thus in a typical epoxidation MnLCl (0.56mmol, salen Mn catalyst) and P<sub>3</sub>NO (2.32mmol) in chlorobenzene (10ml) are added to a 2M solution of NaOCl (26.5mmol) cooled to 0° under N<sub>2</sub>, and stirred at 0° for 15 minutes, followed by simultaneous addition of indene (10ml) and further NaOCl (79.5mmol) (syringe pumps) during 30 minutes. The reaction is complete within 1 hour at 0°, and can be performed on a multi-kilogram scale to provide indene epoxide in 89% yield and with an optical purity of 88% e.e. [Senanayake et al. *Tetrahedron Lett* **37** 3271 1996.]

A **dimeric Jacobsen ligand**, formed by linking two units with a methylene bridge has been prepared and crystallised from CH<sub>2</sub>Cl<sub>2</sub>/pentane. Its **bis-Mn catalyst** exhibited improved retention in a poly-dimethylsiloxane membrane for the asymmetric epoxidation of olefines [Janssen et al. *Tetrahedron: Asymmetry* **8** 3481 1997].

**[(*R*)-(+)-2,2'-Bis(diphenylphosphino)-1,1'-binaphthyl]palladium(II) dichloride ([*R*-BINAP]<sub>2</sub>PdCl<sub>2</sub>)** [115826-95-4] **M 800.0, m 255-260° (dec), [α]<sub>D</sub><sup>24</sup> +682° (c 0.5, CHCl<sub>3</sub>)**. This catalyst is prepared by dissolving a mixture of *R*-BINAP (500mg, 0.803mmol, see [76189-55-4]), η-allyl-η-cyclopentadienyl-Pd(II) (84.2mg, 0.396mmol, see [1271-03-0]), *cf* Tatsuno et al. *Inorg Synth* **19** 220 1979) in \*C<sub>6</sub>H<sub>6</sub> (5ml), in a Schlenk tube

under  $N_2$  (dried over  $P_2O_5$ ), and heated at  $80^\circ$  for 3 hours to give a deep red solution which is then evaporated to dryness. The red solid is then dissolved in  $CH_2Cl_2$  (10ml), filtered and the filtrate is evaporated *in vacuo* down to ca 4ml, diluted with  $Et_2O$  (40ml) and cooled to  $4^\circ$  whereby red-purple crystals of  $[R-BINAP]_2Pd(0)$  separate, which on washing with cold  $Et_2O$  and drying *in vacuo* give analytically pure  $[R-BINAP]_2Pd(0)-Et_2O$  (480mg, 90%). The 1:1 ratio of  $Et_2O$  to Pd is confirmed by  $^1H$  NMR spectroscopy. It has  $^1H$  NMR (90MHz,  $*C_6D_6$ , TMS) with  $\delta_H$  at 5.88 (t,  $J = 7.6Hz$ , 8H), 6.19 (t,  $J = 7.3Hz$ , 4H), 6.8-7.6 (m, 40H), 7.83 (d,  $J = 8.2Hz$ , 4H), 8.36 (brs, 8H); and the  $^{31}P\{^1H\}$  NMR (36MHz,  $*C_6D_6$ , referenced to external 85%  $H_3PO_4$  at  $\delta$  0.0) has  $\delta_P$  at 27.3. This Pd(0) complex catalyses asymmetric arylation in the Heck reaction.  $[R-BINAP]_2PdCl_2$  is obtained by stirring *R*-BINAP (249mg, 0.40mmol) in  $*C_6H_6$  (4.0ml) with  $PdCl_2(MeCN)_2$  (104mg, 0.40mmol) in  $*C_6H_6$  (4.0ml) overnight. The yellow precipitate is collected by filtration, washed with  $*C_6H_6$  and dried *in vacuo*. It is crystallised by dissolving in  $Me_2CO$ , layering with hexane and setting aside at  $\sim 25^\circ$  to give red crystals of  $[R-BINAP]_2PdCl_2$  (238mg, 75%, m 255-260 $^\circ$ ) of analytical purity, and suitable for X-ray crystallographic structure determination which confirmed its structure. [Ozawa et al. *Organometallics* **12** 4188 1993, cf review by Shimizu, Nagasaki & Saito *Tetrahedron* **61** 5405 2005].

**Bis(trifluoromethanesulfonyl)amine metal salts  $[NTf_2]$ , bis-(trifluoromethanesulfonyl)amide metal salts,  $(CF_3SO_2)_2N$ Metal** see [82113-65-3].  $NTf_2$  readily forms salts, e.g. with Li, Ba, Ca, Al, Zn, Al and lanthanides, which act as *Lewis Acids* in effectively catalysing Diels-Alder reactions [Kobayashi et al. *Chem Lett* 307 1995, Handy et al. *Synlett* 565 1995] and acylation reactions [Mikami et al. *Synlett* 171 1996]. It forms an *N*-trimethylsilyl derivative (TMSNTf<sub>2</sub> prepared from allylTMS and  $NTf_2$ , see above) which catalyses Diels-Alder reactions between methyl acrylate and various dienes [Mathieu & Ghosez *Tetrahedron Lett* **38** 5497 1997], as well as Friedel-Crafts acylation reactions of anisole, and allylation and bis-allylation of carbonyl derivatives [Oshii et al. *Synlett* 1145 1997].  $NTf_2$  is a weakly coordinating counter ion which confers stability to catalytic metal complexes such as the gold catalyst  $[Ph_3PAu(I)]^+ [NTf_2]^-$ , see [866395-16-6].

***R*(-)- and *S*(+)- 4,12-Bis(diphenylphosphino)-[2.2]-paracyclophane {*R*- and *S*- PHANEPhos, (*R*)- and (*S*)-5,11-bis(phenylphosphino)-tricyclo[8.2.2.2<sup>4,7</sup>]hexadeca-hexaene}** [*R*] 364732-88-7; [*S*] 192463-40-4] **M** 576.6, **m** 222-226 $^\circ$ , 224-276 $^\circ$ ,  $[\alpha]_D^{25}$  *R* -34 $^\circ$ , *S* +34 $^\circ$  (c 1,  $CHCl_3$ ). These are chiral ligands for the rhodium precatalysts  $(COD)Rh\{(R)-[2.2]PHANEPhos\}^+ OTf^-$  and  $(COD)Rh\{(S)-[2.2]PHANEPhos\}^+ OTf^-$  (see preparation and purification below) which are highly enantioselective in the hydrogenation of dehydroamino acid methyl esters generally, and also of methyl 1-Ac-4-Boc-1,4,5,6-tetrahydropyrazine-1-carboxylate which is difficult to reduce. When they are pre-reduced in MeOH, they form  $(MeOH)_2Rh\{(R)-[2.2]PHANEPhos\}^+ OTf^-$  whose activity is more pronounced and reduction is complete in less than 60 minutes by passing  $H_2$  gas through the MeOH solution at temperatures between  $-45^\circ$  to  $+50^\circ$ , with high substrate to catalyst ratios. The preparations of these ligands (described briefly here for convenience) involve di-bromination of *p*-cyclophane, separation of the *meso*- from the *racemic*-isomers (latter are more soluble in triglyme), reaction of the *rac*-isomer (1 equivalent) with *tert*-BuLi (4.2 equivalents), and  $MgBr_2 \cdot Et_2O$  (2.3 equivalents) in THF then diphenylphosphinyl chloride ( $Ph_2POCl$ , 2.2 equivalents) to give *rac*-4,12-bis(diphenylphosphinyl)-[2.2]-paracyclophane (76% yield). Optical resolution of this racemate (in  $CHCl_3$ ) with (+)-dibenzoyl-D-tartaric acid (in EtOAc) at  $60^\circ$ , then cooling to  $\sim 25^\circ$  over 18 hours, gave a white precipitate from which the *R*(-)-enantiomer (>99.5% ee) is isolated. The absolute configuration was obtained from the X-ray crystallographic structure determination of the *R*-dibenzoyl-D-tartrate salt (for definition of the chiral descriptor notation see R.S. Chan, C.K. Ingold and V. Prelog *Angew Chem Int Edn Engl* **5** 385 1966). From the mother liquors the *S*(+)-enantiomer (>99.5% ee) was isolated by washing the diastereomeric tartrate salt (in  $CHCl_3$ ) with 0.5M aqueous NaOH. They are white solids which are separated by SFC (supercritical fluid chromatography using a 25cm Hewlett Packard Chiracel AD column with 15% MeOH/supercritical  $CO_2$  at 1.0ml/minute) with retention times of 13.0 minutes for the *R*-isomer and 17.8 minutes for the *S*-isomer, and they have  $[\alpha]_D^{25}$  -104 $^\circ$  and  $[\alpha]_D^{25}$  +104 $^\circ$  (c 0.1, EtOH) respectively. When *R*-4,12-bis(diphenylphosphinyl)-[2.2]-paracyclophane (4.0g, 6.58mmol) is treated with  $Cl_3SiH$  (10ml, 99mmol, 10025-78-2) in *p*-xylene (80ml), and slowly heated to  $140^\circ$ , kept at this temperature for 18 hours, a further aliquot of  $Cl_3SiH$  (5ml, 49.5mmol) is added, and heating is continued for 24

hours, the reduction appears complete. The mixture is then cooled to  $-10^{\circ}$ , quenched with 30% aqueous NaOH (60ml, care violent reaction), extracted with EtOAc (3 x 80ml), the extract is dried ( $\text{MgSO}_4$ ), filtered and evaporated to dryness. Degassed MeOH (30ml, by sonication) is added to the residue which is filtered and dried to give the pure diphosphino *R*-[2.2]PHANEPhos as a white solid. It has  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ , with residual  $\text{CHCl}_3$  as reference at  $\delta$  7.27) with  $\delta_{\text{H}}$  at 7.54, (m, 2H), 7.45 (m, 3H), 7.41 (m, 2H), 7.24 (m, 3H), 6.62 (*ortho and meta*, 2H), 6.58 (dd,  $J = 7.6$  and  $5.0\text{Hz}$ , 1H), 3.13 (m, 1H), 3.03 (*ortho and meta*, 2H), 2.65 (m, 1H); and  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ , with residual  $\text{CHCl}_3$  as reference at  $\delta$  77.0) with  $\delta_{\text{C}}$  at 143.2 (d,  $J = 16.1\text{Hz}$ ), 139.6 (d,  $J = 11.2\text{Hz}$ ), 139.2 ( $J = 2.4\text{Hz}$ ), 137.4 (d,  $J = 12.0\text{Hz}$ ), 137.2 (d,  $J = 10.4\text{Hz}$ ), 135.8 (d,  $J = 22.5\text{Hz}$ ), 134.3 (d,  $J = 4.0\text{Hz}$ ), 133.3 (t,  $J = 4.0\text{Hz}$ ), 132.9 (d,  $J = 20.1\text{Hz}$ ), 132.7, 129.4, 128.44, 128.39 (d,  $J = 8.8\text{Hz}$ ), 128.2 (d,  $J = 8.0\text{Hz}$ ), 35.8, 33.2; and  $^{31}\text{P}$  NMR (160MHz,  $\text{CDCl}_3$ , referenced to external 85%  $\text{H}_3\text{PO}_4$  at  $\delta$  0.0) with  $\delta_{\text{P}}$  at  $-0.53$ . The enantiomeric *S*-[2.2]PHANEPhos is similarly obtained.

$\text{Rh}(\text{COD})\{(\text{R})\text{-[2.2]PHANEPhos}\}^+ \text{OTf}^-$  is prepared using Schlenk equipment under  $\text{N}_2$  from the red solution of *R*-[2.2]PHANEPhos (630mg, 1.09mmol) and  $[(\text{Rh COD})_2]^+ \text{OTf}^-$  (512mg, 1.09mmol, see 12092-47-6 for chloride), in  $\text{CH}_2\text{Cl}_2$  (10ml) which is stirred for 30 minutes, evaporated *in vacuo*, MeO-*t*-Bu (10ml) is added to the residue, sonicated for 30 minutes, stirred vigorously for 30 minutes, filtered and the solid is dried under  $\text{N}_2$  to give the orange *R*- precatalyst (860mg, 85%). It has  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ , with residual  $\text{CHCl}_3$  as reference at  $\delta$  7.27) with  $\delta_{\text{H}}$  at 8.58 (m, 2H), 7.85 (m, 1H), 7.31 (m, 2H), 7.59 (m, 1H), 7.43 (m, 1H), 7.36 (m, 2H), 7.19 (m, 2H), 6.55 (br d,  $J = 8.0\text{Hz}$ , 1H), 6.43 (dd,  $J = 8.0$  and  $4.0\text{Hz}$ , 1H), 4.50 (br s, 2H), 2.77 (m, 1H), 2.67 (m, 1H), 2.54 (m, 1H), 2.48 (m, 1H), 2.20 (om, 3H), 2.03 (m, 1H); and  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ , with residual  $\text{CHCl}_3$  as reference at  $\delta$  77.0), due to the non-equivalence of the P atoms, some carbon signals are second order multiplets which are noted as multiplets with  $\delta_{\text{C}}$  at 142.2, 139.9 (m), 139.3 (m), 139.0 (t,  $J = 4.0\text{Hz}$ ), 138.3 (m), 134.6, 133.7, 132.2 (t,  $J = 4.0\text{Hz}$ ), 131.2 (m), 130.8 (m), 130.59, 130.58 (m), 129.2 (t,  $J = 4.2\text{Hz}$ ), 128.9 (t,  $J = 4.4\text{Hz}$ ), 100.5 (dt,  $J = 8.8$  and  $3.2\text{Hz}$ ), 91.6 (dt,  $J = 8.0$  and  $7.0\text{Hz}$ ), 35.1, 34.5, 32.5, 28.8; and  $^{31}\text{P}$  NMR (160MHz,  $\text{CDCl}_3$ , referenced to external 85%  $\text{H}_3\text{PO}_4$  at  $\delta$  0.0) with  $\delta_{\text{P}}$  at 32.7 (d,  $J_{\text{P-Rh}} = 146.1\text{Hz}$ ). The enantiomeric  $\text{Rh}(\text{COD})\{(\text{S})\text{-[2.2]PHANEPhos}\}^+ \text{OTf}^-$  is similarly prepared. [Pye, Rossen et al. *J Am Chem Soc* **119** 6207 1997.] For the catalytic hydrogenation of  $\beta$ -keto esters to  $\beta$ -hydroxy esters using  $\text{Ru}(\text{II})\{[\text{2.2]PHANEPhos}\}(\text{trifluoroacetate})_2$  with up to 96% ee, see Pye, Rossen et al. *Tetrahedron Lett* **39** 4441 1998.

***R*(-)- and *S*(+)- 4,12-Bis[di(3,5-xylyl)phosphino]-[2.2]-paracyclophane {*R*- and *S*- xylyl-PHANEPhos, (*R*- and (*S*)-5,11-bis[di(3,5-xylyl)phenylphosphino]-tricyclo-[8.2.2.2<sup>4,7</sup>]hexadeca-hexaene} [(*R*) 325168-89-6; (*S*) 325168-88-5] M 688.9, m 234-238 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{22}$  *R*  $-61.0^{\circ}$ , *S*  $+61.0$  (c 0.1, EtOH). These chiral ligands are prepared and purified in much the same way as for [2.2]PHANEPhos above except that 3,5-dimethylphenylphosphinyl chloride is used to introduce the phosphorous group into the *p*-cyclophane. When these enantiomers are complexed with ruthenium and chiral ethylenediamine-type bases, they form powerful precatalysts that in the presence of  $\text{H}_2$ , will catalyse the reduction of a range of simple aromatic, heteroaromatic as well as  $\alpha,\beta$ -unsaturated ketones with high ee values. Only one example, ***R*-{[xylyl-PHANEPhos]- $\text{Ru}_2\text{Cl}(\text{S,S})\text{-DPEN}$ }** is described here. It is prepared from the red solution of  $[(\text{C}_6\text{H}_6)\text{RuCl}_2]_2$  (218mg, 0.436mmol, see [37366-09-9]), *R*-xylyl-PHANEPhos (0.60mg, 0.871mmol) in dry DMF (6ml) by heating at  $100^{\circ}$  for 4 hours, to which is then added *S,S*-1,2-diphenylethylene-1,2-diamine (185mg, 0.871mmol, *S,S*-DPEN, [29841-69-8]), and heating, and stirring is continued at  $100^{\circ}$  for 1.5 hours. The mixture is filtered, the filtrate is evaporated *in vacuo*, treated with  $\text{Et}_2\text{O}$  (10ml) and MeOH (100), the yellow catalyst is filtered off, washed with MeOH (12ml) and dried (420mg, 45% yield). It has  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{H}}$  at 1.78 (ddd,  $J = 15.0, 10.5, 4.5\text{Hz}$ , 1H), 2.07 (s, 6H), 2.27 (s, 6H), 2.59 (s, 1H), 2.78 (m, 1H), 4.01 (m, 1H), 4.40 (m, 1H), 4.49 (m, 1H), 6.36 (d,  $J = 8.0\text{Hz}$ ), 6.42 (d,  $J = 8.0\text{Hz}$ ), 6.60 (m, 2H), 6.79 (m, 1H), 6.96 (m, 2H), 7.06 (s, 1H), 7.13 (m, 3H), 8.04 (m, 2H), 8.21 (m, 1H); and  $^{31}\text{P}$  NMR (162MHz,  $\text{CDCl}_3$ , referenced to external 85%  $\text{H}_3\text{PO}_4$  at  $\delta$  0.0) with  $\delta_{\text{P}}$  at  $+46.06$  (s). The crude catalysts performed equally as well as the purified precatalysts. [Burk et al. *Org Lett* **2** 4173 2000.]**

**{1,2-Bis[(4*S*)-4-isopropyl-2-oxazolin-2-yl]benzene}zinc(II) Chloride** [131380-93-3] M436.7, m 202 $^{\circ}$ . The zinc complex is obtained by adding a solution of the free ligand (300mg, 1mmol, see [131380-80-8]) in dry THF (5ml) to a solution of fused anhydrous  $\text{ZnCl}_2$  (136mg, 1mmol, dried by melting in high vacuum and cooling under argon) in dry THF (10ml) and collecting the crystals, but if crystals are not readily formed, evaporate to

dryness and recrystallise the residue from EtOH to give the colourless complex (420mg, 98%). It has IR (KBr) with  $\nu_{\max}$  at 1655 (C=N)  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (300MHz,  $\text{CD}_3\text{CN}$ ) has  $\delta$  at 0.84 (d,  $J = 6.9\text{Hz}$ , 6H,  $\text{CH}(\text{CH}_3)_2$ ), 0.95 (d,  $J = 7.1\text{Hz}$ , 6H,  $\text{CH}(\text{CH}_3)_2$ ), 2.52-2.62 (m, 2H,  $\text{CH}(\text{CH}_3)_2$ ), 4.50-4.62 (m, 6H,  $\text{OCH}_2\text{CHN}$ ), 7.77-7.81 (m, 2H, 4-H and 5-H), 7.85-7.89 (m, 2H, 3-H and 6-H) from TMS; and the  $^{13}\text{C}$  NMR (100MHz,  $\text{CD}_3\text{CN}$ ) has  $\delta$  at 14.85 (q), 18.59 (q), 30.92 (d), 71.08 (d), 71.39 (t), 126.57 (s), 132.25 (d), 134.05(d), 168.99 (s). [Bolm et al. *Chem Ber* **124** 1173 1991.] Like many chiral bis(oxazolin-2-yl) complexes it can be involved in a variety of metal catalysed asymmetric synthesis such as allylation, aziridination, cyclopropanation, Diels-Alder and retro Diels Alder, Mukaiyama aldol condensation and hydrosilylation [see reviews by Ghosh et al. *Tetrahedron: Asymmetry* **9** 1 1998, Pflatz *Acta Chem Scand* **50** 189 1996, Johnson & Evans *Acc Chem Res* **33** 325 2000, Jørgensen et al. *Acc Chem Res* **32** 605 1999.]

**1,3-Bis(2,4,6-trimethylphenyl)-imidazol-2-ylidene-gold(I) chloride (IMesAuCl)** [852445-81-9] **M 537.1, m ~280-286°**. The gold complex is prepared by adding  $\text{Ag}_2\text{O}$  (0.5mmol) to  $\text{IMes}^+ \text{Cl}^-$  (1mmol) in  $\text{CH}_2\text{Cl}_2$  and the suspension is stirred for 3 hours at  $\sim 25^\circ$  whereby it became clear; then a solution of  $\text{Au}(\text{Me}_2\text{S})\text{Cl}$  (1mmol, see [29892-37-3]) in  $\text{CH}_2\text{Cl}_2$  is added dropwise and stirred for a further 4 hours. After filtration through Celite, the filtrate is evaporated to a small volume and hexane is added to allow the complex to separate. It is collected and dried (49% yield). Its  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ , TMS) has  $\delta_{\text{H}}$  at 7.08 (s, 2H), 6.97 (bs, 4H), 2.33 (s, 6H), 2.09 (s, 12H), and its  $^{13}\text{C}$  NMR (100MHz,  $\text{THF}-d_8$ , TMS) has  $\delta_{\text{C}}$  at 134.68, 129.49, 122.12, 21.13, and 17.75 (one signal is missing, coincident signals?). HRMS-FAB Calc for  $\text{C}_{21}\text{H}_{25}\text{AuN}_2\text{Cl}$  ( $\text{M}^+ + \text{H}$ ) is 537.1372; and found 537.1361. It catalyses the intramolecular [4+2] cycloadditions of 1-en-6-yne, or aralkynes with alkenes, in the presence of  $\text{AgSbF}_6$ ; also the related **1-methyl-3-(2,4,6-trimethylphenyl)-imidazol-2-ylgold(I) chloride** is almost equally effective [Nieto-Oberhuber *J Am Chem Soc* **127** 6178 2005].

The saturated 4,5-dihydro- analogue **1,3-bis(2,4,6-trimethylphenyl)-imidazolin-2-ylidene-gold(I) chloride (SMeAuCl)** [852445-82-0] **M 539.1**, behaves in much the same manner as  $\text{IMesAuCl}$ , chemically and catalytically. It has recently been shown that both  $\text{IMesAuCl}$  and  $\text{SMeAuCl}$  can be made by metal exchange from the cheaper respective  $\text{IMesCuCl}$  and  $\text{SMeCuCl}$  (see below) in 71-90% yields. Thus a mixture of  $[\text{AuCl}(\text{SMe}_2)]$  (0.4-0.47mmol, [29892-37-3]) and  $[\text{CuCl}(\text{NHC})]$  (1 equivalent) in  $\text{CH}_2\text{Cl}_2$  (5ml) is stirred at  $40^\circ$  for 1-2 hours, then filtered through Celite and the solvent is removed *in vacuo*.  $\text{CH}_2\text{Cl}_2$  (3ml) and petroleum ether (10ml) are added to the residue, the almost pure Au complex is filtered off, washed with petroleum ether (3 x 5ml) until colourless, and dried *in vacuo*. Their catalytic activities from these different preparations are the same. [Furst & Cazin *Chem Commun* **46** 6924 2010.]

**Chloro[(1,1'-biphenyl-2-yl)di-tert-butylphosphine]gold(I) (JohnPhos-gold(I) chloride, JohnPhos-AuCl)** [854045-93-5] **M 530.8, m 237-240°**. The gold complex is made from sodium tetrachloroaurate(III) dihydrate [13874-02-7] or hydrogen tetrachloroaurate(III) [see 16961-25-4] (1 mmol) in  $\text{H}_2\text{O}$  and the yellow solution is cooled in ice, to which is added bis(2-hydroxyethyl)sulfide [see 111-48-8] (3mmol) dropwise over  $\sim 45$  minutes with stirring. Then a solution of JohnPhos [224311-51-7] (1mmol) in EtOH is added dropwise, whereby the gold complex separates as a white solid which is filtered off (43% yield), washed with MeOH and dried *in vacuo*. The  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 7.87 (td,  $J = 7.7, 1.7\text{Hz}$ , 1H), 7.51 (m, 5H), 7.31 (m, 1H), 7.13 (dd,  $J = 8.0, 1.0\text{Hz}$ , 2H), 1.41 (d,  $J = 15.6\text{Hz}$ , 18H); the  $^{31}\text{P}$  NMR (160MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at 63.11; and the  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 150.16 (d,  $^2J_{(^{13}\text{C},^{31}\text{P})} = 13.5\text{Hz}$ ), 142.10 (d,  $^3J_{(^{13}\text{C},^{31}\text{P})} = 6.3\text{Hz}$ ), 133.46 (d,  $^4J_{(^{13}\text{C},^{31}\text{P})} = 3.0\text{Hz}$ ), 133.22 (d,  $^3J_{(^{13}\text{C},^{31}\text{P})} = 7.5\text{Hz}$ ), 129.19, 128.67, 128.22, 126.67 (d,  $^3J_{(^{13}\text{C},^{31}\text{P})} = 6.8\text{Hz}$ ), 126.06 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 45.3\text{Hz}$ ), 37.75 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 25.7\text{Hz}$ ), 30.84 (d,  $^2J_{(^{13}\text{C},^{31}\text{P})} = 6.7\text{Hz}$ ); HRMS-ESI calc for  $\text{CHAuNP}$  ( $\text{M} + \text{MeCN}\text{-Cl}$ ): 536.1781, found: 536.1779. [Nieto-Oberhuber et al. *J Am Chem Soc* **127** 6168 2005, Al-Sa'Ady et al. *Inorg Synth* 191 1985.]

The acetonitrile complex salt derivative (**acetonitrile**)-(2-biphenyl)-di-tert-butylphosphinogold(I) hexafluoroantimonate [866641-66-9] **M 772.1, has m 198-203°(dec)**. It catalyses a highly efficient intramolecular addition of  $\beta$ -ketoamide to unactivated olefins leading to highly substituted lactams in the presence of  $\text{AgOTf}$  (e.g. *N*-allyl *N*-benzyl benzoylacetamide to *cis*-3-benzoyl-1-benzyl-4-methylpyrrolidin-2-one in 99% yield) [Zhou & Che *J Am Chem Soc* **129** 5828 2007]. JohnPhos-AuCl and related complexes, in the presence of  $\text{AgSbF}_6$  (i.e. the  $\text{SbF}_6^-$  salts), catalyse intramolecular [4+2] cycloadditions of 1,3-enynes or arylalkynes and alkenes with high efficiency [Nieto-Oberhuber et al. *J Am Chem Soc* **127** 6168 2005].

**Chloro[2-dicyclohexyl(2',4',6'-tri-isopropylbiphenyl)phosphine]gold(I) [X-Phos-gold(I) chloride, X-Phos-**

**AuCl]** [854045-94-6] **M 709.1, m 243-250°**. This complex is prepared in much the same way as JohnPhos-AuCl but using a solution containing 1mmol of X-Phos [564483-18-7] in EtOH instead and giving a 61% yield of the X-Phos-AuCl as a white solid. The  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 7.57-7.55 (m, 2H), 7.23 (m, 1H), 7.05 (s, 2H), 2.95 (hept,  $J = 7.0\text{Hz}$ , 1H), 2.21 (hept,  $J = 6.7\text{Hz}$ , 2H), 2.05-2.01 (m, 4H), 1.83-1.71 (m, 6H), 1.67-1.61 (m, 2H), 1.53-1.42 (m, 3H), 1.35 (s, 3H), 1.34 (s, 3H), 1.28 (s, 3H), 1.26 (s, 3H), 1.23-1.13 (m, 7H), 0.92 (s, 3H), 0.91 (s, 3H); the  $^{31}\text{P}$  NMR (160MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at 38.48; and the  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 150.6, 145.48, 133.82 (d,  $^3J_{(^{13}\text{C},^{31}\text{P})} = 8.1\text{Hz}$ ), 132.04 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 3.2\text{Hz}$ ), 130.35 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 1.8\text{Hz}$ ), 127.08 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 7.1\text{Hz}$ ), 121.70, 37.23 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 33.5\text{Hz}$ ), 34.30, 30.93, 30.81, 27.01 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 13.5\text{Hz}$ ), 26.69 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 12.9\text{Hz}$ ), 24.36, 23.21 HRMS-ESI calc for  $\text{C}_{33}\text{H}_{49}\text{AuPNa}$  ( $\text{M}^+ + \text{Na}$ ): 731.2823, found: 731.2789. [Nieto-Oberhuber et al. *J Am Chem Soc* **127** 6168 2005, Al-Sa'Ady et al. *Inorg Synth* 191 1985.] Its catalytic activities are similar to those of JohnPhos-AuCl above.

**Chloro(triphenylphosphine)gold(I)** [14243-64-2] **M 495.7, m 240-241°, 242-243°**.  $\text{Ph}_3\text{PAuCl}$  is prepared under argon by adding a solution of hydrogen tetrachloroaurate (2.0g, 5.87mmol,  $\text{H}[\text{AuCl}_4] \cdot n\text{H}_2\text{O}$ ,  $n \sim 3$ , "brown gold chloride" see [16903-35-8]) in EtOH (10ml), which has been filtered to remove any insoluble Au-containing material, to a stirred solution of  $\text{Ph}_3\text{P}$  (3.08g, 11.7mmol) in absolute EtOH (50ml of 100%, warm if necessary to dissolve). The solution immediately becomes colourless and after stirring for 15 minutes the white microcrystalline solid is collected (fritted glass filter, porosity 4), washed with EtOH (2 x 5ml), then  $\text{Et}_2\text{O}$  (3 x 15ml), and dried *in vacuo* (0.1mm) to give the analytically pure complex (1.9-2.3g 67-80%, or up to 92% if equipment is thoroughly dried, solvents degassed, and working in a strict inert atmosphere). It (~1.g) crystallises from  $\text{CH}_2\text{Cl}_2$  (10ml) by slow addition of *n*-pentane (60ml) and cooling to  $-25^\circ$  to give small slender white needles within a few minutes, and it also crystallises from  $\text{CHCl}_3$  solution by addition of heptane. [Braunstein et al. *Inorg Synth* **27** 218 1990, Bruce et al. *Inorg Synth* **26** 324 1986, cf JohnPhos-AuCl above].  $\text{Ph}_3\text{PAuCl}$  has also been prepared by adding  $\text{Ph}_3\text{P}$  (1.5g) in  $\text{Me}_2\text{CO}$  (5ml) to a solution of  $\text{HAuCl}_4$  (1g, 0.5mol equivalents, "chlorogold(III) acid" see [16903-35-8]) in  $\text{Me}_2\text{CO}$  at  $20^\circ$ . The volume of the solution is reduced to *ca* 5ml and the complex that separates (1.2g, 83%) is collected and dried or crystallised as before [Burgess et al. *JCS, Dalton Trans* 1161 1983].

*Alternatively*, to the orange solution of sodium tetrachloroaurate(III). $2\text{H}_2\text{O}$  (1mmol, [13874-02-7]) dissolved in  $2\text{H}_2\text{O}$ , cooled in ice, is slowly added 2,2'-thiodiethanol {3mmol, i.e. bis(2-hydroxyethyl)sulfide [see 111-48-8]} at  $0^\circ$  during ~45 minutes, followed dropwise by a solution of the  $\text{Ph}_3\text{P}$  ligand (1mmol, or other  $\text{PX}_3$  ligand) in EtOH which causes a white solid to separate. This is filtered off, washed with MeOH, and dried *in vacuo* (0.1mm) to give  $\text{Ph}_3\text{PAuCl}$  (or other  $\text{PX}_3\text{AuCl}$ ) as a white solid in ~88% yield. It is recrystallised as described above. [Li et al. *J Org Chem* **73** 4323 2008,]. The  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 2.03-1.93 (m, 9H), 1.89-1.85 (m, 6H), 1.75-1.72 (m, 3H), 1.52-1.41 (m, 6H), 1.35-1.21 (m, 9H); the  $^{31}\text{P}$  NMR (160MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at 57.12 (34.224ppm reported also at 300MHz,  $\text{CDCl}_3$ , from external  $\text{H}_3\text{PO}_4$ ); and  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 33.28 (d,  $^1J_{(^{13}\text{C},^{31}\text{P})} = 31\text{Hz}$ ), 30/75, 26.96 (d,  $^2J_{(^{13}\text{C},^{31}\text{P})} = 12.2\text{Hz}$ ), 25.80 (d,  $^3J_{(^{13}\text{C},^{31}\text{P})} = 1.5\text{Hz}$ ), HRMS-ESI calc for  $\text{C}_{20}\text{H}_{36}\text{AuPNa}$  ( $\text{M}^+ + \text{MeCN}\cdot\text{Cl}$ ): 518.2251, found: 518.2241. [Nieto-Oberhuber et al. *J Am Chem Soc* **127** 6168 2005, Al-Sa'Ady et al. *Inorg Synth* 191 1985, Li et al. *J Org Chem* **73** 4323 2008.] Its catalytic activities are similar to those of JohnPhos-AuCl above. It has also been used in the cyclisation of *O*-propargyl carbamates to alkylideneoxazolidiones at room temperature, *via* a 5-exo-digonal pathway [Ritter et al. *Synlett* 3309 2006]. It is soluble in  $\text{Me}_3\text{CO}$ ,  $\text{CH}_2\text{Cl}_2$ ,  $\text{CHCl}_3$ ,  $^*\text{C}_6\text{H}_6$  and THF, but not in EtOH or hexane, and is used for the preparation of a variety of  $\text{Ph}_3\text{PAu}$  salts and cluster complexes. It is stable indefinitely in air, and note that  $^{31}\text{P}$  NMR studies of  $\text{Ph}_3\text{PAuCl}$  showed no change in the chemical shift (at 33.8ppm in THF) after bubbling  $\text{O}_2$  through the solution of 24 hours, indicating how stable it is compared with cationic  $\text{Ph}_3\text{PAu}^+$  species in the presence of  $\text{O}_2$  [Liu et al. *J Am Chem Soc* **128** 11332 2006]. X-Ray crystallography showed that the molecule of  $\text{Ph}_3\text{PAuCl}$  is linear with a P-Au-Cl angle of  $179.68^\circ$  and Au-P distance of  $2.235\text{\AA}$  [Baezinger et al. *Acta Cryst* **B32** 962 1976].

**Chlorotris(triphenylphosphine)rhodium(I) (Wilkinson's catalyst)** [14694-95-2] **M 925.2, m 138°(dec), 140°(dec), 157-158°(dec)**. It is prepared by adding  $\text{PPh}_3$  (12g, recrystallised from EtOH to remove any  $\text{PPh}_3\text{O}$ )

in hot EtOH (350ml) to  $\text{RhCl}_3 \cdot 3\text{H}_2\text{O}$  (2g) in 95% EtOH (70ml) in a flask purged with  $\text{N}_2$ , refluxed in the presence of  $\text{N}_2$  for 2 hours, and the red solid that is formed is collected on a sintered glass funnel and washed with small volumes of dry  $\text{Et}_2\text{O}$  (50ml) to give the catalyst (6.25g, 88%). It forms dark burgundy crystals from hot EtOH after refluxing for 30 minutes. When the solution is heated for only 5 minutes, orange crystals are formed. Heating the orange crystals in EtOH yields red crystals. Crystallisation from  $\text{Me}_2\text{CO}$  gives the orange crystals. The dimorphic forms have similar IR spectra, but the X-ray diffraction patterns are slightly different. Excess of  $\text{PPh}_3$  can be recovered by adding  $\text{H}_2\text{O}$  to the EtOH filtrates from recrystallisation and, collecting it after standing for 2-3 days. [Osborne et al. *J Chem Soc (A)* 1711 1966, Osborne & Wilkinson *Inorg Synth* **10** 67 1967, Bennett & Donaldson *Inorg Chem* **16** 655 1977.] The solubilities are as follows: in  $\text{CH}_2\text{Cl}_2$  ~2% (25°), in toluene 0.2% (25°), and less soluble in  $\text{Me}_2\text{CO}$ , MeOH, BuOH and AcOH, but insoluble in petroleum ethers and cyclohexane. It reacts with donor solvents such as pyridine, DMSO and MeCN. Solutions of  $\text{RhCl}[(\text{PPh}_3)_3]$  are air sensitive, absorb  $\text{O}_2$ , and as well as the solid should be stored under  $\text{O}_2$ -free  $\text{N}_2$  or argon. On heating solutions in  $^*\text{C}_6\text{H}_6$ , toluene or methyl ethyl ketone the chlorine-bridged *dimer*  $[(\text{PPh}_3)_3\text{RhCl}_2\text{Rh}(\text{PPh}_3)_3]$  is formed as salmon-red crystals which absorb oxygen even in the solid state, but can be converted to the *monomer* in refluxing EtOH containing  $\text{PPh}_3$ . Red solution of  $\text{RhCl}[(\text{PPh}_3)_3]$  absorb  $\text{H}_2$  reversibly at 760mm/25°, become yellow in colour, and are effective in catalysing the reduction of C-C double bonds and triple bonds, often in a selective manner. [Osborn & Wilkinson *Inorg Synth* **28** 77 1990.] The stoichiometric and catalytic reactions of  $\text{RhCl}[(\text{PPh}_3)_3]$  have been reviewed [Jardine *Prog Inorg Chem* **28** 63 1981]. In the presence of 1.5 mol of  $\text{Et}_2\text{Zn}$ , 1 mol of 4-phenylbut-3-ene-2-one (and related  $\alpha$ -ene-ones) reacts with 1.7mol of  $\text{CF}_3\text{Cl}$  in THF at 5° to give 4-phenyl-3-(trifluoromethyl)butan-2-one (and related ketones) in moderate yields in the presence of catalytic amounts of  $\text{RhCl}[(\text{PPh}_3)_3]$  (~0.02mol) [Sato, Omote, Ado and Kumadaki *Org Synth* **83** 177 2006].

**Chromium (III) acetylacetonate**  $[\text{Cr}(\text{acac})_3]$  [21679-31-2] **M 349.3, m 212-216°, 216°, b 340°, pK<sup>25</sup> 4.0** (see **chromic chloride**). Purify it by dissolving 6g in hot  $^*\text{C}_6\text{H}_6$  (20ml) and adding petroleum ether (75ml) slowly. Cool to room temperature then chill on ice, filter off and dry in air to give 2.9g. It also crystallises from EtOH. It is soluble in heptane,  $^*\text{C}_6\text{H}_6$ , toluene and pentane-2,4-dione at 20-40°. It forms a 1:2 complex with  $\text{CHCl}_3$ . [Fernelius & Blanch *Inorg Synth* **V** 130 1957, Steinbach & Burns *J Am Chem Soc* **80** 1839 1958, Beilstein **1** H 782, **1** II 836, **1** IV 3673.]  $\text{Cr}(\text{acac})_3$  catalyses the oxidation of methacrylic acid esters in the presence of  $\text{H}_2\text{O}_2$  which provides a new route to pyruvic esters [Inoue et al. *Chem Lett (Jpn)* **18** 99 1989]. **TOXIC** — potential carcinogen.

**Cobalt Oxazoline Palladacycles (COPs)** are organocobalt-palladium complexes which catalyse the asymmetric rearrangements of non-chiral allylic trichloroacetamidates with very high enantiomeric selectivity (>90%) to provide chiral allylic amines [it is an aza-Claisen rearrangement, '**The Overman Rearrangement**' Overman & Carpenter *Org React* **66** 2005, Kirsch, Overman and Watson *J Org Chem* **69** 8101 2004]; and in the presence of phenols stereospecific cross-coupling also occurs to provide chiral phenoxyallyl ethers with very high (>90%) enantiomeric selectivity [Kirsch, Overman and White *Org Lett* **9** 911 2007, Overman & Carpenter *Org React* **66** 2005].

**(S)-(+)-COP-OAc dimer catalyst**  $\{S(\text{COP-OAc})_2$ ; **di- $\mu$ -acetobis** $[\eta^5$ -**(S)**-**(pR)**-2-(2'-(4'-methyleneethyl)-oxazoliny)l)cyclopentadienyl, 1-C-3'-N)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt} **dipalladium**  $\}$  [222400-03-5] **M 1512-2, m 189-194° (250-257° dec.),  $[\alpha]_D^{24} +942^\circ$  (c 0.215,  $\text{CHCl}_3$ )**. This COP is prepared in a Schlenk flask flushed with argon containing ( $\eta^5$ -**(S)**-2-(4'-methyleneethyl)-oxazoliny)l)cyclopentadienyl-( $\eta^4$ -tetraphenylcyclobutadiene)cobalt (9.6g, 16.2mmol, see 22240-02-4) in glacial acetic acid (96ml) to which is added palladium(II) acetate (3.6g, 16.2mmol, must be recrystallised from  $^*\text{C}_6\text{H}_6$ ), and the red solution is heated at 95° for 30 minutes when an orange precipitate separates. After cooling to ~25°, the solid is collected, washed with glacial acetic acid (50ml) and dried *in vacuo* to give almost analytically pure **S**-(COP-Ac)<sub>2</sub> (8.9g, 73%) as mustard coloured crystals. It has IR (thin film) with  $\nu_{\text{max}}$  at 3061, 2961, 1583 (C=N), 1501, 1417 (acetate bridge), 1366, 1181, 1069  $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (500MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{H}}$  at -0.01 (d,  $J = 6.6\text{Hz}$ , 3H,  $\text{CH}_3$ ), 0.46 (d,  $J = 7.1\text{Hz}$ , 3H,  $\text{CH}_3$ ), 1.72-1.81 (m, 1H, CH), 1.96 (s, 3H,  $\text{CH}_3$ ), 2.98 (td,  $J = 9.0, 3.1\text{Hz}$ , 1H, CH), 3.36 (t,  $J = 9.0\text{Hz}$ , 1H,  $\text{CH}_2$ ), 4.08 (dd,  $J = 8.6, 3.9\text{Hz}$ , 1H,  $\text{CH}_2$ ), 4.23 (t,  $J = 2.4\text{Hz}$ , 1H, CH), 4.62 (d,  $J = 1.4\text{Hz}$ , 1H, CH), 4.68 (d,  $J = 2.0\text{Hz}$ , 1H,

CH), 7.20-7.29 (m, 12H, ArH), 7.64 (m, 8H, ArH); and the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 13.3, 18.8, 24.2, 29.1, 65.0, 71.2, 76.2, 79.4, 85.0, 85.4, 86.8, 98.1, 126.2, 128.1, 129.5, 136.1, 170.8, and 180.9.

The **(R)-(-)-COP-OAc dimer enantiomer** [849592-74-1] **M 1512.2, m 241-251<sup>o</sup> dec.** is prepared in the same way but using the enantiomeric starting material. [Stevens & Richards *Organometallics* **18** 1346 1999, Anderson et al. *Org Synth* **84** 148 2007.] The reactions of trichloroacetimidate derivatives of *Z*-2-alken-1-ols with phenolic nucleophiles in the presence of chiral COP-OAc dimer catalysts yield 3-aryloxy-1-alkenes in high yield (63-90%) and high enantiomeric purity (90-97% ee); and are compatible with the presence of base-labile substituents in either reactant [Kirsch, Overman and White *Org Lett* **9** 911 2007].

**(S)-(+)-COP-Cl dimer catalyst** {**S(COP-Cl)<sub>2</sub>**; **di- $\mu$ -chlorobis[ $\eta^5$ -(*S*)-(p*R*)-2-(2'-(4'-methylethyl)-oxazoliny)cyclopentadienyl, 1-C-3'-N)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt[dipalladium]**} [581093-92-7] **M 1464.98, m 204-205<sup>o</sup>, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +1201<sup>o</sup> (c 0.1,  $\text{CH}_2\text{Cl}_2$ ).** It is obtained from the above *S*-(COP-OAc)<sub>2</sub> (8.9g, 5.9mmol) in acetone (59ml) by stirring vigorously at ~25<sup>o</sup> with aqueous NaCl (2M, 30ml) to give a homogeneous mixture until the reaction is complete (~4 hours, checked periodically by sampling a filtered aliquot and examining the  $^1\text{H}$  NMR spectra). The yellow complex which separates is filtered off, washed with H<sub>2</sub>O (125ml) then Me<sub>2</sub>CO (20ml) and dried *in vacuo* to give almost pure (COP-Cl)<sub>2</sub> (8.2g, 95%) as a mustard coloured solid. It can be purified further by filtering a solution in CH<sub>2</sub>Cl<sub>2</sub> through a short plug of Celite and eluting with CH<sub>2</sub>Cl<sub>2</sub> to give an analytically pure orange solid (97% recovery) which exists as a 1.0:0.7 mixture of dimers. It has IR (thin film) with  $\nu_{\text{max}}$  at 3061, 2961, 1602, 1502, 1370, 1185 cm<sup>-1</sup>; the  $^1\text{H}$  NMR (500MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 0.70-0.74 (m, 6H, CH<sub>3</sub>), 0.76 (d, *J* = 6.8Hz, 3H, CH<sub>3</sub>), 0.80 (d, *J* = 7.0Hz, 3H, CH<sub>3</sub>), 2.20-2.33 (m, 2H, CH), 3.03-3.15 (m, 2H, CH), 3.34 (t, *J* = 9.0Hz, 1H, CH<sub>2</sub>), 3.43 (t, *J* = 8.9Hz, 1H, CH<sub>2</sub>), 4.16-4.22 (m, 2H, CH<sub>2</sub>), 4.28 (t, *J* = 2.5Hz, 1H, CH), 4.40 (t, *J* = 2.5Hz, 1H, CH), 4.70 (d, *J* = 2.0Hz, 1H, CH), 4.73 (d, *J* = 2.0Hz, 1H, CH), 4.98 (d, *J* = 1.5Hz, 1H, CH), 4.99 (d, *J* = 1.5Hz, 1H, CH), 7.16-7.22 (m, 12H, ArH), 7.23-7.30 (m, 12H, ArH), 7.58-7.62 (m, 8H, ArH), 7.66-7.71 (m, 8H, ArH); and the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 14.3, 14.5, 19.0 (2C), 29.0, 29.2, 65.6, 66.0, 71.4, 71.5, 76.5, 76.8, 81.0, 84.2, 84.5, 84.8, 84.9, 85.7, 87.1, 98.6, 98.7, 126.4, 126.5, 128.2, 128.3, 129.4, 129.5, 135.5, 135.6, 171.1, 171.2. The enantiomeric purity was determined by conversion into the acetylacetonate derivative **COP-acac** (see below) and the enantiopurity was shown to be >99% by HPLC [Diacel Chiralpak AD-H (0.46cm x 25cm) column, with 1.0ml/minute flow rate and eluted with 95:5 hexanes:*iso*-PrOH]. The retention time for *S*-COP-acac is 5.1 minutes whereas that of *R*-COP-acac would have been ~3.1 minutes [Anderson, Kirsch and Overman *Org Synth* **84** 148 2007].

**(R)-(-)-COP-Cl dimer catalyst** {**R(COP-Cl)<sub>2</sub>**; **di- $\mu$ -chlorobis[ $\eta^5$ -(*R*)-(p*R*)-2-(2'-(4'-methylethyl)-oxazoliny)cyclopentadienyl, 1-C-3'-N)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt[dipalladium]**} [612065-00-6] **M 1464.98, m 205-208<sup>o</sup>, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +1240<sup>o</sup> (c 0.1,  $\text{CH}_2\text{Cl}_2$ )** is obtained and purified as for its *S*-enantiomer above but using the *R*-enantiomeric intermediates. These planar chiral COP-Cl dimers catalyse the [3+3]-sigmatropic rearrangements of a wide range of non-chiral *E*-allylic trichloroacetimidates into the corresponding transposed chiral allylic trichloroacetamides (from which the free base can be obtained) in high yields and very high enantioselectivities [Anderson & Overman *J Am Chem Soc* **125** 12412 2003, Anderson, Overman and Watson *Org Synth* **82** 134 2005]. The reactions of trichloroacetimidate derivatives of *Z*-2-alken-1-ols with phenolic nucleophiles in the presence of chiral COP-Cl dimer catalysts yield 3-aryloxy-1-alkenes in high yield (63-90%) and high enantiomeric purity (90-97% ee); and are compatible with the presence of base-labile substituents in either reactant [Kirsch, Overman and White *Org Lett* **9** 911 2007].

**(S)-(+)-COP-acac monomer catalyst** {**acetylacetonato[ $\eta^5$ -(*S*)-(p*R*)-2-(2'-(4'-methylethyl)-oxazoliny)cyclopentadienyl, 1-C-3'-N)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt[palladium]**} [805315-09-7] **M 796.1, m 100-104<sup>o</sup>(dec.), [ $\alpha$ ]<sub>D</sub><sup>28</sup> +246.1<sup>o</sup>, [ $\alpha$ ]<sub>577</sub><sup>28</sup> +250.2<sup>o</sup>, [ $\alpha$ ]<sub>546</sub><sup>28</sup> +175.5<sup>o</sup>, [ $\alpha$ ]<sub>435</sub><sup>28</sup> -59.8<sup>o</sup>, [ $\alpha$ ]<sub>405</sub><sup>28</sup> -91.4<sup>o</sup> (c 1.00,  $\text{CHCl}_3$ ).** *S*-(+)-COP-acac is obtained as a monomer from the above dimer *S*-(COP-OAc)<sub>2</sub> (1.0g, 0.66mmol), sodium acetylacetonate (954mg, 6.8mmol, see [15435-71-9]), Me<sub>2</sub>CO (6.6ml) and H<sub>2</sub>O (3.3ml) by stirring vigorously for 24 hours at ~25<sup>o</sup>. The mixture is then extracted with CH<sub>2</sub>Cl<sub>2</sub> (10ml), the extract is separated, dried (MgSO<sub>4</sub>), filtered and concentrated to give almost pure *S*-(+)-COP-acac (1.0g, quantitative) as an orange solid. It can be further purified by filtering through a silica-gel column and eluting with *iso*-PrOH:hexanes (5:95) and evaporating the orange band to give an analytically pure complex. See above *S*-(COP-Cl)<sub>2</sub> for HPLC data; and it has  $R_{\text{F}}$  0.59 (silica gel, hexanes-EtOAc, 80:20). It has IR (thin film) with  $\nu_{\text{max}}$  at 3058, 2960, 1597, 1579, 1508.

1399, 1265, 1183, 1067  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 7.64 (d,  $J = 7.6\text{Hz}$ , 8H), 7.20-7.30 (m, 12H), 5.26 (s, 1H), 5.17 (s, 1H), 4.89 (d,  $J = 1.6\text{Hz}$ , 1H), 4.47 (s, 1H), 4.27 (dd,  $J = 8.4, 5.6\text{Hz}$ , 1H), 3.67 (t,  $J = 9.0\text{Hz}$ , 1H), 3.29-3.31 (m, 1H), 2.26-2.27 (m, 1H), 2.00 (s, 3H), 1.94 (s, 3H), 0.84 (d,  $J = 7.6\text{Hz}$ , 3H), 0.82 (d,  $J = 8.0\text{Hz}$ , 3H); and the  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 14.8, 18.6, 27.2, 27.6, 29.0, 65.1, 71.6, 76.0, 80.0, 85.1, 85.8, 86.3, 99.4, 126.0, 127.8, 128.1, 129.2, 136.0, 172.0, 184.9, 186.3. [Kirsch et al. *J Org Chem* **69** 8101 2004, Anderson, Kirsh and Overman *Org Synth* **84** 148 2007].

The isomer **(R)-(-)-COP-acac monomer catalyst {acetylacetonato[ $\eta^5$ -(R)-(pR)-2-(2'-(4'-methyl-ethyl)-oxazoliny)cyclopentadienyl, 1-C-3'-N-( $\eta^4$ -tetraphenylcyclobutadiene)cobalt]palladium}** [CASRN is the same as preceding *S*-(+)-enantiomer] **M 796.1, m 100-104 $^\circ$ (dec),  $[\alpha]_{\text{D}}^{28} -246.1^\circ$ ,  $[\alpha]_{577}^{28} -250.2^\circ$ ,  $[\alpha]_{546}^{28} -175.5^\circ$ ,  $[\alpha]_{435}^{28} +59.8^\circ$ ,  $[\alpha]_{405}^{28} +91.4^\circ$  (c 1.00,  $\text{CHCl}_3$ )** can be obtained and purified as for its *S*-enantiomer above but using the *R*-enantiomeric intermediates. These monomeric COPs catalyse the asymmetric rearrangements of non-chiral allylic trichloro-acetimidates into chiral transposed allylic trichloroactamides in good yield with high asymmetric induction (>90% ee). These are more soluble catalysts than the COP dimers, and because they are soluble in a much wider variety of solvents, the reactions can be carried out at high substrate concentration (e.g. at ~2.6M) [Kirsch et al. *J Org Chem* **69** 8101 2004]. The reactions of trichloroacetimidate derivatives of *Z*-2-alken-1-ols with phenolic nucleophiles in the presence of chiral COP-acac monomer catalysts yield 3-aryloxy-1-alkenes in high yield (63-90%) and high enantiomeric purity (90-97% ee); and are compatible with the presence of base-labile substituents in either reactant [Kirsch, Overman and White *Org Lett* **9** 911 2007].

**(S)-(+)-COP-hfacac {hexafluoroacetylacetonato[ $\eta^5$ -(S)-(pR)-2-(2'-(4'-methyl-ethyl)oxazoliny)cyclopentadienyl, 1-C-3'-N-( $\eta^4$ -tetraphenylcyclobutadiene)cobalt]palladium}** [805315-08-6] **M 904.0, m 108-110 $^\circ$ (dec.)  $[\alpha]_{\text{D}}^{28} +271.2^\circ$ ,  $[\alpha]_{577}^{28} +277.6^\circ$ ,  $[\alpha]_{546}^{28} +214.8^\circ$ ,  $[\alpha]_{435}^{28} -57.7^\circ$ ,  $[\alpha]_{405}^{28} -86.8^\circ$  (c 1.02,  $\text{CHCl}_3$ )**. *(S)*-(+)-COP-hfacac is obtained as a monomer from the above dimer *S*-(COP-OAc)<sub>2</sub> (1.0g, 0.66mmol), sodium hexafluoroacetylacetonate (1.6g, 6.8mmol, see [22466-49-5]),  $\text{Me}_2\text{CO}$  (6.6ml) and  $\text{H}_2\text{O}$  (3.3ml) by stirring vigorously at ~25 $^\circ$  for 9-11 hours (check periodically by sampling a filtered aliquot and examining the  $^1\text{H}$  NMR spectra until completion). The solid that separates is filtered off, washed with  $\text{H}_2\text{O}$  (10-20ml) and dried *in vacuo* over  $\text{P}_2\text{O}_5$  to give *(S)*-(+)-COP-hfacac monomer (1.1g, 91%). A pure sample can be obtained by filtration through a short plug of silica gel with  $\text{CH}_2\text{Cl}_2$  as eluent (in an 89% yield). It has  $R_{\text{F}}$  0.63 (silica gel, hexanes-EtOAc, 80:20). It has IR (thin film) with  $\nu_{\text{max}}$  at 3061, 2964, 1629, 1598, 1509, 1475, 1258, 1208, 1150  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (500MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 7.54-7.56 (m, 8H, ArH), 7.19-7.29 (m, 12H, ArH), 5.95 (s, 1H, CH), 4.68 (d,  $J = 2.3\text{Hz}$ , 1H, CH), 4.90 (d,  $J = 2.3$ , 1H, CH), 4.53 (t,  $J = 2.3\text{Hz}$ , 1H, CH), 4.33 ( $J = 8.6, 5.3\text{Hz}$ , 1H,  $\text{CH}_2$ ), 3.72 (t,  $J = 9.4\text{Hz}$ , 1H,  $\text{CH}_2$ ), 3.45 (td,  $J = 9.4, 5.1\text{Hz}$ , 1H, CH), 2.05-2.08 (m, 1H, CH), 0.83 (d,  $J = 7.0\text{Hz}$ , 3H,  $\text{CH}_3$ ), 0.79 (d,  $J = 6.9\text{Hz}$ , 3H,  $\text{CH}_3$ ); and the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 14.9, 18.4, 29.3, 65.3, 72.3, 76.7, 78.9, 84.1, 84.8, 87.7, 90.1, 97.6, 116.6, 118.8, 126.5, 127.9, 128.9, 135.5, 173.3, 173.6, 174.1. [Kirsch et al. *J Org Chem* **69** 8101 2004, Anderson et al. *Org Synth* **84** 148 2007].

Enantiomer **(R)-(-)-COP-hfacac {hexafluoroacetylacetonato[ $\eta^5$ -(R)-(pR)-2-(2'-(4'-methyl-ethyl)-oxazoliny)cyclopentadienyl, 1-C-3'-N-( $\eta^4$ -tetraphenylcyclobutadiene)cobalt]palladium}** [CASRN is the same as preceding *S*-(+)-enantiomer] **M 904.0, m 108-110 $^\circ$ (dec.)  $[\alpha]_{\text{D}}^{28} -271.2^\circ$ ,  $[\alpha]_{577}^{28} -277.6^\circ$ ,  $[\alpha]_{546}^{28} -214.8^\circ$ ,  $[\alpha]_{435}^{28} +57.7^\circ$ ,  $[\alpha]_{405}^{28} +86.8^\circ$  (c 1.02,  $\text{CHCl}_3$ )** can be obtained and purified as for its *S*-enantiomer above but using the *R*-enantiomeric intermediates. These COP-hfacac complexes, like the related COP-acac complexes, are monomeric and more soluble than the dimeric COPs above. They can be used in a variety of solvents, e.g. THF, MeCN, which allow catalytic asymmetric allylacetimidate rearrangements to be performed at high concentrations and using as little as 1 mole% of catalyst. [Anderson Kirsch and Overman *Org Synth* **84** 148 2007, Kirsch, Overman and Watson *J Org Chem* **69** 8101 2004]. The reactions of trichloroacetimidate derivatives of *Z*-2-alken-1-ols with phenolic nucleophiles in the presence of chiral COP-acac catalysts yield 3-aryloxy-1-alkenes in high yield (63-90%) and high enantiomeric purity (90-97% ee); and are compatible with the presence of base-labile substituents in either reactant [Kirsch, Overman and White *Org Lett* **9** 911 2007].

**Copper(I) [1,3-bis(2,6-di-*iso*-propylphenyl)imidazol-2-ylidene] chloride [1,3-bis(2,6-di-*iso*-propylphenyl)imidazol-2-ylidene copper(I) chloride, (IPr)=CuCl** [578743-87-0] **M 487.59, m >300 $^\circ$** . This NHC (N-heterocyclic carbene) precatalyst is readily prepared in analytical purity in a Schlenk flask and under an argon atmosphere. Dry THF (100ml) is added to a mixture of freshly prepared Cu(I)Cl (1.0g, 10.1mmol), 1,3-bis(2,6-

di-*iso*-propylphenyl)imidazolium chloride (4.29g, 10.1mmol, IPr.HCl see [250285-32-6]) and Na *tert*-butoxide (0.97g, 10.1mmol), and stirred mechanically under argon at  $\sim 25^\circ$  for 20 hours. The mixture is then filtered through a plug of Celite, and the filtrate is evaporated *in vacuo* to give the complex as a white powder (4.59g, 9.40mmol, 94%),  $m > 300^\circ$ . Suitable crystals were grown from  $\text{CH}_2\text{Cl}_2$ /hexane solutions and subjected to single-crystal diffraction studies. The molecular structure was as expected with the imidazole C2, Cu and Cl atoms in the same plane, with the C2-Cu bond length of 1.953Å, the Cu-Cl bond length 2.089Å, and the C2—Cu—Cl angle of  $180.00^\circ$ , i.e. is linear. [Kaur et al. *Organometallics* **23** 1157 1160]. It has IR (neat on a DiComp probe) with  $\nu_{\text{max}}$  at 2968, 2927, 2871, 1680, 1647, 1470, 1457, 809, 764, and 743  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (500MHz,  $\text{Me}_2\text{CO}-d_6$ , downfield from TMS) has  $\delta_{\text{H}}$  at 7.26 (s, 2H), 7.57-7.54 (dd,  $J = 15.5$  and  $8.0\text{Hz}$ , 2H), 7.43-7.41 (d,  $J = 8.0\text{Hz}$ , 4H), 2.70-2.65 (m, 4H), 1.32-1.31 (d,  $J = 7.0\text{Hz}$ , 12H), 1.27-1.25 (d,  $J = 6.5\text{Hz}$ , 12H); the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$  with 77.23 ppm of  $\text{CDCl}_3$  reference) has  $\delta_{\text{C}}$  at 180.2, 146.0, 135.1, 130.6, 124.2, 24.4 and 23.3; and the HRMS (EI.  $m/z$ ) has: found 389.2962, and calculated for  $\text{C}_{14}\text{H}_{16}\text{O}$ : 389.2951 (M - CuCl) [Jurkauskas et al. *Org Lett* **5** 2417 2003.]

(IPr)=CuCl catalyses the conjugate reduction of  $\alpha,\beta$ -unsaturated carbonyl compounds with catalytic amounts of the Cu complex, Na *tert*-BuO and poly(methylhydrosiloxane) as the stoichiometric reductant in  $>81\%$  yields [Jurkauskas et al. *Org Lett* **5** 2417 2003], it catalyses the methenylation of a variety of aliphatic and aromatic aldehydes and ketones in the presence of trimethylsilyldiazomethane ( $\text{TMSCHN}_2$ ),  $\text{Ph}_3\text{P}$  and 2-propanol efficiently, and is a cheaper alternative to the corresponding rhodium catalyst [LeBel et al. *J Org Chem* **72** 144 2007], it catalyses the efficient hydrosilylation of ketones with  $\text{Et}_3\text{SiH}$  to yield the corresponding saturated triethylsilylethers [Kaur et al. *Organometallics* **23** 1157 1160], it catalyses the transfer of the ‘:CHCO<sub>2</sub>Et’ group (from ethyl diazoacetate) to saturated and unsaturated substrates, e.g. olefins, amines or alcohols, in very high yields [Fructos et al. *J Am Chem Soc* **126** 10846 2004], and has been used for the total synthesis of (-)-Angelastin A by catalysing the aziridiation (using  $\text{TsN}=\text{IPh}$ ) [55962-05-5] of a difficult rather electron-deficient cyclopentene intermediate [Trost & Dong *J Amer Chem Soc* **128** 6054 2006].

**NOTE:** More recently it has been shown that the above  $\text{IPr}=\text{CuCl}$  ( $\text{IPrCuCl}$ ) and its 4,5-dihydro-derivative  $\text{SIPrCuCl}$ , are readily prepared by refluxing  $\sim 1.9$ - $2.3\text{mmol}$  of 1,3-bis-(2,6-diisopropylphenyl)imidazolium chloride ( $\text{IPrCl}$ , see [250285-32-6]) or 1,3-bis-(2,6-diisopropylphenyl)imidazolium chloride ( $\text{SIPrCl}$ , see [25278-25-0]) with  $\text{Cu}_2\text{O}$  ( $\sim 1.4$ - $1.5\text{mmol}$ ,  $\sim 95\%$  purity, *ca* 0.65 equivalents) for 24 hours, then excess of  $\text{Cu}_2\text{O}$  is removed to provide the following respective yields in the solvents indicated:  $\text{IPrCuCl}$  (78% from toluene, 88% conversion; 94% from  $\text{H}_2\text{O}$  with 96% conversion; 74% from dioxan) or  $\text{SIPrCuCl}$  (88% from toluene, 99% conversion; 72% from  $\text{H}_2\text{O}$ , 78% conversion). [Cazin and coworkers: Citadella et al. *Dalton Trans* **39** 4489 2010, Son and coworkers: Chun et al. *Organometallics* **29** 1518 2010].

Similarly the analogous [1,3-bis(2,4,6-trimethylphenyl)imidazol-2-ylidene copper(I) chloride, [(IMes)=CuCl]] and the 4,5-dihydro derivative [1,3-bis(2,4,6-trimethylphenyl)imidazol-2-ylidene copper(I) chloride [(SIMes)=CuCl, S is for saturated] can be prepared from the respective *IMesCl* or *SIMesCl* with 0.65 equivalents of  $\text{Cu}_2\text{O}$  by refluxing in toluene for 24 hours (in 86% and 71% yields respectively, with 100% conversion), and by refluxing in  $\text{H}_2\text{O}$  for 24 hours (in 98% and 99% yields respectively, with 100% conversion). [Cazin and coworkers: Citadella et al. *Dalton Trans* **39** 4489 2010]. These are generally slightly less catalytically specific than their IPr counter-parts because the latter have slightly greater steric hindrance.

**Copper(I) bis-[1,3-bis(2,6-di-*iso*-propylphenyl)imidazol-2-ylidene] tetrafluoroborate [(IPr)<sub>2</sub>Cu<sup>+</sup> BF<sub>4</sub><sup>-</sup>] [886061-48-9] **M 927.5, m >300°**. This copper catalyst is prepared in an oven dried septum screw capped vial (in a glove box) by mixing tetrakis(acetonitrile)copper (I) tetrafluoroborate (186mg, 0.5mmol), IPr.Cl (157mg, 0.5mmol, see [250285-32-6]) and *ter*-BuONa (96mg, 1mmol), sealing, then injecting THF (10ml) and stirring outside the box at  $\sim 25^\circ$  for 6 hours. The mixture is filtered through a plug of Celite (THF), and the Cu complex salt is precipitated by adding excess hexane, filtered off, and dried *in vacuo* to give the pure white tetrafluoroborate (452mg, 92%). It has  $^1\text{H}$  NMR (400MHz,  $\text{Me}_2\text{CO}-d_6$ ) with  $\delta$  at 7.54 (t containing a singlet at 7.55,  $J = 7.5\text{Hz}$ , 8H, *p*-CH + NCH), 7.25 (d,  $J = 7.5\text{Hz}$ , 8H, *m*-CH), 2.40 (septet,  $J = 6.8\text{Hz}$ , 8H,  $\text{CH}(\text{CH}_3)_2$ ), 1.04 (d,  $J = 6.8\text{Hz}$ , 24H,  $\text{CH}_3$ ), 0.94 (d,  $J = 6.8\text{Hz}$ , 24H,  $\text{CH}_3$ ); and its  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 177.4 (N-C-N), 145.0 (CH aromatic), 134.7 (C aromatic), 130.6 (C aromatic), 125.3 (=CH-N), 124.5 (CH aromatic), 28.6 ( $\text{CH}(\text{CH}_3)_2$ ), 24.3 ( $\text{CH}_3$ ), 24.2 ( $\text{CH}_3$ ). The hexafluorophosphate salt is similarly prepared in 96% yield starting from tetrakis(acetonitrile)copper (I) hexafluorophosphate [64443-05-6]. The structure of these salts was shown by X-ray crystal analysis (crystals grown from  $\text{CH}_2\text{Cl}_2$ -methyl *tert*-butyl ether solution) to consist of two *N*-heterocyclic carbenes attached to one Cu atom.**

These salts are air- and moisture- stable and are highly active catalysts for the hydroxysilylation of ketones with varying steric constraints, aldehydes and epolisable aldehydes, and esters in THF at  $\sim 25^\circ$ , in the presence of  $\text{Et}_3\text{SiH}$  and *ter*-BuONa in  $>90\%$  yields [Díez-González et al. *Organometallics* **25** 2355 2006].

**Diacetato[(*R*)-(+)- and (*S*)-(–)- 2,2'-Bis(diphenylphospheno)-1,1'-binaphthyl]ruthenium(II) [*R*- and *S*-(BINAP)<sub>2</sub> Ru(II) (OAc)<sub>2</sub>] [(*R*) 132071-87-5; (*S*) 261948-85-0, 104713-03-3] M 841.8, m 185-187°(dec.), 186-188°(dec.), 188-190°(dec.), *R*-  $\lambda_{\text{max}}$  335nm ( $\Delta\epsilon/\text{mol}^{-1}\cdot\text{dcm}^3\cdot\text{cm}^{-1}$  +4.5) and  $\lambda_{\text{max}}$  450nm ( $\Delta\epsilon/\text{mol}^{-1}\cdot\text{dcm}^3\cdot\text{cm}^{-1}$  +2.5); *S*-  $\lambda_{\text{max}}$  335nm ( $\Delta\epsilon/\text{mol}^{-1}\cdot\text{dcm}^3\cdot\text{cm}^{-1}$  -4.5) and  $\lambda_{\text{max}}$  450nm ( $\Delta\epsilon/\text{mol}^{-1}\cdot\text{dcm}^3\cdot\text{cm}^{-1}$  -2.5); ( $c$   $6.5 \times 10^{-5}$  M,  $\text{CH}_2\text{Cl}_2$ ). Typically, the *S*-enantiomer is prepared in a Schlenk tube under dry argon, by dissolving (*S*)-BINAP (1.37g, 2.20mmol, see [76189-56-5]) and  $[\text{RuCl}_2(\text{COD})]_n$  (0.56g, 2.0mmol, calc based on monomer weight, see [50982-12-2]) in dry toluene (55ml) containing  $\text{Et}_2\text{NH}$  (1.2ml, 8.6mmol), and the brown suspension is refluxed with stirring for 12 hours. The clear reddish brown solution is cooled to  $\sim 25^\circ$ , the solvent is removed *in vacuo* and the brown solid residue is dissolved in  $\text{CH}_2\text{Cl}_2$  (40ml), filtered through a Celite pad and the filtrate is evaporated *in vacuo*. A solution of anhydrous NaOAc (0.88g, 11mmol) in *t*-BuOH (107ml) is added to the residue, stirred under reflux for 12 hours in an argon atmosphere, the solvent is removed *in vacuo* and the residue is extracted with absolute  $\text{Et}_2\text{O}$  ( $3 \times 20\text{ml}$ ) and evaporated again *in vacuo*. The brown residue is extracted several times with absolute EtOH (70ml total) and evaporated *in vacuo* to give a yellow-brown solid (1.58g) which is recrystallised from toluene (12ml)/hexane (30ml) mixture to afford purer complex (1.23g, 68% yield). This sample is pure enough for providing good catalytic activity for asymmetric hydrogenations. Analytical purity (to give 0.9g, 50%, 188-190° dec.) can be achieved by further recrystallisation from toluene (14ml)/hexane (13ml), and another amount (0.3g, 17%) can be recovered from the mother liquors by evaporation and recrystallisation of the residue (from 8ml of toluene and 10ml of hexane). Crystals of analytical material were subjected to Xray structure determination that gave the required absolute configuration of the complex, and the circular dichroism spectrum of it and its enantiomer were consistent with the Xray structure. It has IR ( $\text{CH}_2\text{Cl}_2$ ) with  $\nu_{\text{max}}$  at 1452, 1518  $\text{cm}^{-1}$ ; and  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ , TMS) with  $\delta_{\text{H}}$  at 1.80 (s, 2 OCOCH<sub>3</sub>), 6.47-7.84 (m, aromatic H); and  $^{13}\text{C}\{^1\text{H}\}$  NMR (100MHz,  $\text{CDCl}_3$ , TMS) with  $\delta$  at 23.5 (OCOCH<sub>3</sub>), 125.2-138.3 (aromatic C), 188.8 (OCOCH<sub>3</sub>). The *R*-enantiomer can be prepared and purified in the same way.**

The enantiomers of this complex catalyse the asymmetric hydrogenation of pro-chiral ketones and olefins in high yields and ee values. [Ohta, Takaya and Noyori *Inorg Chem* **27** 566 1988, Kumbobyashi et al. *Synlett* 1055 2001.] RhBINAP complexes with ligands other than acetate, e.g.  $\text{Cl}_2$ ,  $^*\text{C}_6\text{H}_6$ ,  $(\text{OCO-}t\text{-Bu})_2$ ,  $(\text{OCOCMe}=\text{CHMe})_2$  also catalyse asymmetric hydrogenations, isomerisations and asymmetric Heck reactions [Shimizu, Nagasaki and Saito *Tetrahedron* **61** 5405 2005.] An excellent asymmetric hydrogenation catalyst for  $\alpha$ -(acylimino)acrylic acids to give the corresponding chiral  $\alpha$ -aminoacylpropionic acids with high ee-ratios is (*S*)-[(BINAP)Rh<sup>+</sup>(OMe)<sub>2</sub>] ClO<sub>4</sub><sup>−</sup>. It is obtained (together with norbornane) when a solution of (*S*)-[(BINAP)Rh<sup>+</sup>(norbornadiene)] ClO<sub>4</sub><sup>−</sup> in MeOH is exposed to H<sub>2</sub> at room temperature. Exactly 2mols are absorbed, and it is isolated as deep red prisms [whose  $^1\text{H}$  NMR ( $\text{CD}_2\text{Cl}_2$ , TMS) has  $\delta_{\text{H}}$  at 7.50 (m) and 6.82 (BINAP), 3.42 (s, CH<sub>3</sub>O); and  $^{31}\text{P}$ NMR ( $\text{CD}_3\text{OD}$  with 5% H<sub>3</sub>PO<sub>4</sub> in  $\text{CD}_3\text{OD}$  as external standard) has  $\delta$  at 53.1 (d,  $J_{\text{Rh-P}} = 206\text{Hz}$ )]. It loses MeOH *in vacuo* to give an equally active catalyst. [Myashita et al. *J Am Chem Soc* **102** 7932 1980.]

**Dichloro(1,5-cyclooctadiene)palladium(II) [Pd(COD)Cl<sub>2</sub>] [12107-56-1] M 285.5, m 210° (dec).** This air-stable catalyst can be prepared by adding 1,5-cyclooctadiene (3ml, 2.2mol, [111-78-4]) to a cooled and filtered solution of PdCl<sub>2</sub> (2.0g, 1 mol) and concentrated HCl (150ml, warm) in EtOH (150ml), whereby the yellow complex separates immediately. After keeping for 10 minutes the solid is filtered off, washed with  $\text{Et}_2\text{O}$  ( $3 \times 30\text{ml}$ ) and dried *in vacuo* (3.1g, 96%). The yellow powder is recrystallised by boiling in  $\text{CH}_2\text{Cl}_2$  (200ml), and evaporating until the onset of crystals. **Dibromo(1,5-cyclooctadiene)palladium(II) [Pd(COD)Br<sub>2</sub>] [12145-47-0] M 374.4, m 213° (dec),** is obtained from the dichloride and NaBr in Me<sub>2</sub>CO, or as above from PdCl<sub>2</sub>/HCl in the presence of NaBr before adding the diene. [Drew & Doyle *Inorg Synthesis*, Coll Vol **28** 348 1990, **13** 52 1972.] *Alternatively*, a solution of sodium chloropalladite tetrahydrate (2g) in MeOH (75ml) and the diene (2ml) can be used, and the pale orange-yellow complex is collected after 1 hour and is recrystallised from glacial AcOH. [Pd(COD)Br<sub>2</sub>] can also be prepared from the dichloride (0.45g) and LiBr (0.1g) in Me<sub>2</sub>CO (20ml), boiled under

reflux for 2 hours, filtered and the filtrate is evaporated to dryness in a vacuum (15mm), the residue is washed with H<sub>2</sub>O, dried (0.22g) and is recrystallised from AcOH (orange-red needles). [Chatt et al. *J Chem Soc* 3413 1957.] The *dichloro-complex* is sparingly soluble in cold EtOH and \*C<sub>6</sub>H<sub>6</sub>, but is more soluble in hot \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub>, Me<sub>2</sub>CO, MeEtCO, (EtO)<sub>2</sub>CO, tetrahydrothiophene 1,1-dioxide (sulfolane) and nitrobenzene. It reacts with Me<sub>2</sub>SO to form *Pd(II)Cl<sub>2</sub>(Me<sub>2</sub>SO)<sub>2</sub>*. The IR (Nujol) has  $\nu_{\max}$  at 1489, 1419, 1337, 1088, 999, 867, 825, 794, 768 325 and 295 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  at 2.69 (CH<sub>2</sub> protons) and 6.32 (CH=CH protons). The *dibromo-complex* is similarly air-stable and has similar solubility in solvents, and its IR (Nujol) has  $\nu_{\max}$  at 1472, 1417, 1333, 1172, 1083, 992, 905, 864, 823, 787, 764, 678, 310, 265, 221, 213, 178 and 126 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  at 2.60 (CH<sub>2</sub> protons) and 6.32 (CH=CH protons). [Beilstein 5 IV 404.]

**Dichloro(1,5-cyclooctadiene)platinum(II) [Pt(COD)Cl<sub>2</sub>]\*** [12080-32-9] **M 374.2, m 220-278° (dec), 285° (dec).** The colourless air-stable dichloro-complex is obtained by adding 1,5-cyclooctadiene\* (6ml) to a warm solution (75°) of hydrated chloroplatinic acid (5g) in AcOH (15ml), stirring, and cooling to ~25° then diluting with H<sub>2</sub>O (50ml), and the black suspension is kept at ~25° for 1 hour. The black solid is collected, washed with H<sub>2</sub>O (50ml), Et<sub>2</sub>O (100ml), then suspended in CH<sub>2</sub>Cl<sub>2</sub> (400ml), heated to boiling and kept at this temperature for 5 minutes. The solution is cooled, mixed with chromatographic grade silica gel (5g) and allowed to settle. The supernatant should be colourless, otherwise portions (~1g) of silica gel should be added until the supernatant is colourless. The solid is filtered off, washed with CH<sub>2</sub>Cl<sub>2</sub> (2 x 50ml) and the combined CH<sub>2</sub>Cl<sub>2</sub> solutions (~500ml) are evaporated until crystallisation (75ml) occurs. The hot CH<sub>2</sub>Cl<sub>2</sub> solution is then poured into petroleum ether (b 60-70°) producing a fine white powder which is collected, washed with petroleum ether (50ml) and dried (2.55g, 80%). Recrystallisation in the same way (dissolving in boiling CH<sub>2</sub>Cl<sub>2</sub> and evaporating till crystallisation sets in) yields white macroscopic crystals. It also crystallises from AcOH (charcoal) as white needles. Its solubilities in solvents are similar to those of Pd(COD)Cl<sub>2</sub> above, i.e. being insoluble in most organic solvents except boiling CH<sub>2</sub>Cl<sub>2</sub>, CHCl<sub>3</sub> and AcOH, and decomposes slowly in Me<sub>2</sub>SO (cf related Pd compound above). Its IR (Nujol) has  $\nu_{\max}$  at 1334, 1179, 1333, 1009, 1083, 871, 834, and 782 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  at 2.71 (CH<sub>2</sub> protons) and 5.62 (*J*<sub>Pt-H</sub> = 65Hz, CH=CH protons). [Drew & Doyle *Inorg Synthesis*, Coll Vol 28 346 1990, 13 47 1972.]

**Dibromo(1,5-cyclooctadiene)platinum(II) [Pt(COD)Br<sub>2</sub>]** [12145-48-1] **M 463.1, m 220-278° (dec)** can be obtained from sodium chloroplatinate hydrate (4g), diene (4ml) and LiBr (2g) in EtOH (80-100ml), kept at ~20° for 2 days, and the precipitate is crystallised from AcOH (charcoal) in pale yellow needles. It is slightly more soluble than the chloro-analogue in organic solvents. [Chatt et al. *J Chem Soc* 2496 1957, Beilstein 5 IV 404.]

\* *Dichloro(1,3,5,7-cyclooctatetraene)Platinum(II)*, *dichloro(dicyclopentadiene)Platinum(II)* and *dichloro(2,5-norbornadiene)Platinum(II)* have been prepared successfully by this procedure and using the respective olefins.

**Dichloro[(1,2,3,6,7,10,11,12- $\eta$  .  $\bar{1}$ )dodecatriene-1,12-diyl]ruthenium(IV), [(dichloro-2,6,10-dodecatrienylene-ruthenium(IV)]** [12170-97-7] **M 334.2, decomposes >200° with part melting at 220°.** This complex is much more stable than the corresponding Ni complex, and is prepared by bubbling butadiene for 7 hours through a solution of RuCl<sub>3</sub> · 3H<sub>2</sub>O (0.43g, [13815-94-6]) in 2-methoxyethanol (25ml) at 90° in a hot water bath. The deep red solution deposits yellow-brown prisms when cooled. These are filtered off and recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/petroleum ether (b 40-60°) to give the analytically pure complex (0.30g, 54.5%) as orange crystals. The IR shows a medium intensity band at  $\nu$  = 1522 cm<sup>-1</sup> assigned to the *trans* C=C co-ordinated to the Ru atom; the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>/TMS) has  $\delta$  at ~5.4 (m, w = 65Hz, 6H, vinylic and non terminal allylic H), 4.9 (d, *J* = 7.5Hz, 2H, terminal allylic H), 3.76 (d, *J* = 11.2Hz, 2H, terminal allylic H), 3.1 (m, w = 20Hz, 4H, methylene H) and 2.37 (m, w = 35Hz, 4H, methylene H); and the measured molecular weight (osmometry in 1.45% \*C<sub>6</sub>H<sub>6</sub> solution) is 340. [Nicholson & Shaw *J Chem Soc (A)* 807 1966.] The <sup>1</sup>H NMR is consistent with a Ru atom enveloped by dodeca-2,6,10-triene and co-ordinated to the two terminal allylic double bonds, i.e. butadiene trimerised about the metal atom, and this structure is confirmed by an X-Ray determination. This showed that Ru is bipyramidal, with the Cl atoms at the apices of the pyramids with three co-ordinate bonds (to the allylic and the C6=C7 double bonds) in the central plane. [Lydon et al. *Proc Chem Soc* 421 1964.]

The *dichloro-dodecatriene-ruthenium* complex is a highly efficient catalyst for the *one-pot* internal redox process that converts allylic alcohols (RC=C-C(OH)) into carbonyl compounds (RC-C-C=O, ketones or aldehydes) in the presence of CsCO<sub>3</sub> under N<sub>2</sub> with high yields and turnover frequencies (TOF) in THF or H<sub>2</sub>O [Cadierno et al. *Chem Commun* 232 2004, see also van der Drift et al. *J Organomet Chem* 650 1 2002]. It is also an efficient catalyst for the de-protection of *N*-allylic and *N*-diallylic substrates (with ~3mol% of Ru) in aqueous solution at

90° under N<sub>2</sub> in very high yields (typically >95%) to give the free amine and propionaldehyde (allyl product); and can be carried out on a preparative scale [Cadierno et al. *Chem Commun* 4086 2005].

**Dichloro(η<sup>5</sup>-pentamethylcyclopentadienyl)iridium(III) dimer {di-μ-chloro-dichlorobis(η<sup>5</sup>-pentamethylcyclopentadienyl)diridium (III), [Rh(η<sup>5</sup>-C<sub>5</sub>Me<sub>5</sub>)Cl<sub>2</sub>]<sub>2</sub>}** [12354-84-6] **M 796.7, m >230°**. The complex is prepared by stirring IrCl<sub>3</sub>·xH<sub>2</sub>O (10g, 26mmol), pentamethylcyclopentadiene (5g, 36mmol, Cp' see [4045-44-7]) and MeOH (300ml), purging the apparatus with N<sub>2</sub> for 5 minutes then refluxing the solution also under N<sub>2</sub> for 48 hours. After cooling to room temperature the orange crystals are filtered off in air through a glass sinter, the filtrate is concentrated *in vacuo* to ~50ml and cooled to give a second crop of crystals which are combined with the first crop, washed with Et<sub>2</sub>O (3 x 50ml) and air dried. The *Ir-dimer* (10.7g, 85%) can be obtained as an orange microcrystalline solid in analytical purity by recrystallisation using the minimum volume of CHCl<sub>3</sub> to dissolve it, filtering off insoluble material if present, and adding slowly twice that volume of hexane. The complex is stable in air at room temperature without obvious decomposition for several years. It is soluble in chlorinated solvents, much less so in Me<sub>2</sub>CO, EtOH or hydrocarbon solvents, but is somewhat soluble in H<sub>2</sub>O. The halogen atoms undergo metathesis and can be replaced by PF<sub>6</sub> and MeCN to form complexes such as [Ir(η<sup>5</sup>-C<sub>5</sub>Me<sub>5</sub>)(NCMe)<sub>3</sub>][PF<sub>6</sub>]<sub>2</sub>. Its reactions can be studied by following the <sup>1</sup>H NMR (CDCl<sub>3</sub>) signal of the single C<sub>5</sub>Me<sub>5</sub> resonance at δ 1.73s. [White et al. *Inorg Synth* **29** 229 1992.] The *Ir-dimer* directly catalyses the C3 alkylation of a variety of indoles using aliphatic and benzylic alcohols in >80% yields [Grigg et al. *Tetrahedron* **65** 4375 2009]. It is used as a precursor of N-3-substituted-N-1-pyrimidyl-imidazolium-Ir-pentamethylcyclopentadiene which catalyses transfer hydrogenation from isoPrOH to ketones, e.g. cyclohexanone, and imines, e.g. benzylideneaniline, to form the respective alcohols and amines [Gnanamagri et al. *Organometallics* **28** 321 2009]. It has also been used as a precursor to efficient phosphine-free Ir-Cp' chiral catalysts, e.g. with *N*-sulfonyl *S,S*-1,2-diphenylethylenediamine, for the hydrogenation of 2-substituted quinolines to produce chiral 2-substituted 1,2,3,4-tetrahydroquinolines in >90% yields with up to 99% ee [Li et al. *Org Lett* **20** 5265 2008].

**Dichloro(η<sup>5</sup>-pentamethylcyclopentadienyl)rhodium(III) dimer {di-μ-chloro-dichlorobis(η<sup>5</sup>-pentamethylcyclopentadienyl)dirhodium (III), [Rh(η<sup>5</sup>-C<sub>5</sub>Me<sub>5</sub>)Cl<sub>2</sub>]<sub>2</sub>}** [12354-85-7] **M 618.1, m >230°**. The complex is readily prepared by stirring RhCl<sub>3</sub>·3H<sub>2</sub>O (10g, 42mmol), pentamethylcyclopentadiene (6g, 44mmol, Cp' see [4045-44-7]) and MeOH (300ml), purging the apparatus with N<sub>2</sub> for 5 minutes then refluxing the solution also under N<sub>2</sub> for 48 hours. After cooling to room temperature the dark red crystals are filtered off in air through a glass sinter, the filtrate is concentrated *in vacuo* to ~50ml and cooled to give a second crop of crystals which are combined with the first crop, washed with Et<sub>2</sub>O (3 x 50ml) and air dried to give the *Rh-dimer* (11.25g, 95%). It can be obtained in analytical purity by recrystallisation using the minimum volume of CHCl<sub>3</sub> to dissolve it, filtering off insoluble material if present, and adding slowly twice that volume of hexane. The complex is stable in air at room temperature without obvious decomposition for several years. It is soluble in chlorinated solvents, much less so in Me<sub>2</sub>CO, EtOH or hydrocarbon solvents, but is somewhat soluble in H<sub>2</sub>O. The halogen atoms undergo metathesis and can be replaced by PF<sub>6</sub> and MeCN to form complexes such as [Rh(η<sup>5</sup>-C<sub>5</sub>Me<sub>5</sub>)(NCMe)<sub>3</sub>][PF<sub>6</sub>]<sub>2</sub>. Its reactions can be studied by following the <sup>1</sup>H NMR (CDCl<sub>3</sub>) signal of the single C<sub>5</sub>Me<sub>5</sub> resonance at δ 1.60s. [White et al. *Inorg Synth* **29** 229 1992.] In the presence of base the rhodium dimer is a good catalyst for the hydrogenation of olefins [Gill et al. *J Chem Soc, Dalton Trans* 617 1978], for hydrosilylation of olefins [Millan et al. *J Mol Catal* **26** 89 1984, Millan et al. *J Chem Soc, Chem Commun* 673 1981], and for the disproportionation of aldehydes [Cook et al. *J Chem Soc, Dalton Trans* 2342 1981].

**Dicobalt octacarbonyl (cobalt carbonyl)** [10210-68-1] **M 341.9, m 51°, 51-52° (dec), d 1.87**. Co<sub>2</sub>(CO)<sub>8</sub> has been identified as the cobalt catalyst in hydroformylation, hydrogenation and homologation reactions involving H<sub>2</sub> and CO (synthesis gas). The earlier preparation [Glimont & Blanchard *Inorg Synth* **2** 238 1946] has been improved due to the availability of steel pressure vessels and CO under pressure. Thus Raney Co (4 to 6g) was placed with Et<sub>2</sub>O (145ml) under CO at 3200psi in a steel pressure bomb and heated with shaking for 5 or 6 hours at 150°. The pressure during the heating process dropped from 4900 to 4300psi and then to 2200psi when the bomb was cooled to ~25°. The insoluble material was removed by centrifugation, and the reddish-brown clear solution (160ml, which includes washings and transfers) contained 8.9g of Co<sub>2</sub>(CO)<sub>8</sub>. It is kept in a closed pressure-adjusted Pyrex bottle and aliquots siphoned as required because it decomposes slowly to liberate CO in an open vessel and should be stored in the cold. The solubility of the catalyst in Et<sub>2</sub>O at ambient temperature is 7.5g per 100ml of solution. Cobalt on Kieselguhr "Co-100 powder" which contains 12 to 15% of Co can be substituted for Raney Co. [Adkins & Kresk *J Am Chem Soc* **70** 383 1948.] The above preparation can also be

carried out in  $^*C_6H_6$  which has the advantage of being less volatile than  $Et_2O$ , and in which reactions could be carried out at lower temperatures. It has been kept at a concentration of  $1.0 \times 10^{-2}$  M in 50ml of  $^*C_6H_6$ . [Adkins & Kresk *J Am Chem Soc* **71** 3051 1949.] It forms orange-brown *air-sensitive* crystals on recrystallisation from *n*-hexane under a carbon monoxide atmosphere. It has been sublimed *in vacuo* (orange platelets), and is available commercially as a solid moistened with 5-10% of hexane. It is insoluble in  $H_2O$ , but soluble in organic solvents such as alcohols,  $Et_2O$ ,  $^*C_6H_6$ , CS, slowly decomposed by HCl, and  $H_2SO_4$ , but more rapidly by  $HNO_3$ . [Wender et al. *J Am Chem Soc* **71** 4160 1949, Ojima et al. *J Am Chem Soc* **109** 7714 1987; see also Hileman in *Preparative Inorganic Reactions*, Ed. Jolly, Vol 1 p 101 1987].

Carbon monoxide is **VERY POISONOUS** as it complexes with haemoglobin. Great care should be exercised when using the catalyst as well as CO, and all operations should be carried out in an efficient fume cupboard. The **TOXIC LEVEL** of CO is 50ppm ( $\sim 55\text{mg}/\text{m}^3$ ). **The ANTIDOTE should be at hand and always available in laboratories using CO; and staff should be trained to administer it.** On completion of reactions the autoclave should be filled with  $H_2$  at 1000-1300psi and heated at 110-135° for 45 minutes where cobalt compounds are transformed to the metal. If the product of reaction does not contain oxidisable compounds, the mixture is heated in air on a steam bath when copious evolution of CO occurs and the metal is deposited as a cobalt mirror. A third procedure to decompose  $Co_2(CO)_8$  and related CO compounds is to shake or stir the mixture with dilute  $H_2SO_4$  until evolution of gas ceases. [Wender et al. *J Am Chem Soc* **72** 4375 1950.]

**2,6-Di-iso-propylphenylimino-neophylidene[(S)-(-)-BIPHEN]molybdenum(VI)** **{(S)-Schrock-Hoveyda Catalyst; Mo(N-2,6-di-iso-Pr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(CHCMe<sub>2</sub>Ph)(syn-(S)-(-)-tert-Bu<sub>2</sub>Me<sub>4</sub>(BIPHEN); molybdenum, [(S- or R-)** 3,3'-bis(1,1'-dimethylethyl)-5,5',6,6'-tetramethyl[1,1'-biphenyl]-2,2'-diolato(2-)-κO,κO'} [2,6-bis(1-methylethyl)benzenaminato(2-)](2-methyl-2-phenylpropylidene)- } [S- 205815-80-1; R- 329735-77-5; RS-300344-02-9] **M 755.9.** The key intermediate is **2,6-di-iso-propylphenylimino-neophylidene-bistriflate-dimethoxyethane-molybdenum (VI) complex** {molybdenum, [2,6-bis(1-methylethyl)benzenaminato(2-)] [1,2-dimethoxy-κO]ethane} [2-methyl-2-phenylpropylidene]bis(1,1,1-trifluoromethanesulfonato-κO)-; [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(=CHMePh)-(OTf)<sub>2</sub>-DME] [126949-63-1] which is prepared in three steps. **Firstly** (preferred method): A solution of  $MoO_2Cl_2(THF)_2$  (10.2g, 29.0mmol, see [13637-68-8, 556907-19-8; 12081-12-8]) in DME (200ml) at -30° is stirred vigorously while the following are added sequentially (i) a solution of 2,6-lutidine (12.4g, 116mmol) in DME (10ml at -30°) during 3-5 minutes, (ii) a solution of  $Me_3SiCl$  (31.5g, 290mmol, see [75-77-4]) in DME (40ml, at -30°) over a period of 3-5 minutes, and (iii) a solution of 2,6-di-iso-propylaniline (10.3g, 58mmol, see [24544-04-5]) in DME (15ml at -30°) during 15 minutes. The colour of the solution alters from pale yellow to bright red-orange to deep red-orange as a solid separated. The mixture is stirred as it warmed to room temperature during 6 hours. The mixture is then heated to 50° for 5 hours, filtered through Celite, to remove 2,6-lutidine hydrochloride, which is washed with  $Et_2O$  until the washings are clear; all filtrates are combined and evaporated *in vacuo* to give *bis(2,6-di-iso-propylphenylimino)dichloro-dimethoxyethane molybdenum* [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)<sub>2</sub>Cl<sub>2</sub>DME] (16.7g, 27.5 mmol, 95%) as a brick red solid which is almost of analytical purity (elemental C, H, N and Cl for  $MoC_{28}H_{44}Cl_2N_2O_2$ ), and used in the next step. Its <sup>1</sup>H NMR (400MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 7.01 (d, 4, arom H<sub>m</sub>), 6.89 (t, 2, arom H<sub>p</sub>), 4.29 (sept, 4, CHMe<sub>2</sub>), 3.44 (s, 6, MeOCH<sub>2</sub>CH<sub>2</sub>OMe), 3.18 (s, 4, MeOCH<sub>2</sub>CH<sub>2</sub>OMe), 1.25 (d, 24, CHMe<sub>2</sub>) and the <sup>13</sup>C NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 154.1 (arom C<sub>ipso</sub>), 145.5 (arom C<sub>p</sub>), 128.0 (arom C<sub>m</sub>), 123.9 (arom C<sub>o</sub>), 71.3 (MeOCH<sub>2</sub>CH<sub>2</sub>OMe), 62.8 (MeOCH<sub>2</sub>CH<sub>2</sub>OMe), 28.2 (CHMe<sub>2</sub>), 25.2 (CHMe<sub>2</sub>). **Secondly:** Neophyl magnesium chloride (100ml, 0.5M, 50mmol, see [35293-35-7]) in  $Et_2O$  is added dropwise to a stirred solution of the preceding complex [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)<sub>2</sub>Cl<sub>2</sub>DME] (15.1g, 25mmol) in  $Et_2O$  (250ml) at -30° (all the complex need not have dissolved). The colour of the solution alters from red to orange as  $MgCl_2$  separates, the mixture is allowed to thaw to 25° and is stirred at this temperature for 3 hours. The mixture is filtered through Celite, the filtrate is concentrated *in vacuo*, kept at -40° to provide orange crystals (usually three crops) of almost analytically pure (elemental C, H and N for  $MoC_{44}H_{60}N_2$ ) *bis(2,6-di-iso-propylphenylimino)-bis(neophyl) molybdenum* [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)<sub>2</sub>(-CH<sub>2</sub>C Me<sub>2</sub>Ph)<sub>2</sub>] (14.3g, 80%). Its <sup>1</sup>H NMR (400MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 7.45 (d, 4H, aromatic), 7.24 (t, 4H, aromatic), 7.10 (t, 2H, aromatic), 7.00-6.92 (m, 6H, aromatic), 3.65 (sept, 4H, CHMe<sub>2</sub>), 1.72 (s, 4H, CH<sub>2</sub>CMe<sub>2</sub>Ph), 1.49 (s, 12H, CH<sub>2</sub>CMe<sub>2</sub>Ph), 1.11 (d, 24H, CHMe<sub>2</sub>) and the <sup>13</sup>C NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 153.3, 142.7, 128.6, 126.8, 126.4, 126.0, 123.2, 78.6, 39.5, 32.0, 27.9, 23.5. **Thirdly:** A solution of triflic acid (4.42g, 29.45mmol) in DME (15ml) is added slowly to an orange solution of the preceding

Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)<sub>2</sub>(-CH<sub>2</sub>CMe<sub>2</sub>Ph)<sub>2</sub> (14.2g, 20mmol) in DME (150ml) at -30°, then allowed to thaw to ~25°, stirred overnight, and the deep yellow solution is evaporated to dryness *in vacuo*. The yellow residue is extracted with chilled (~0°) toluene (100ml), filtered, evaporated *in vacuo*, and the dark yellow residue is recrystallised from Et<sub>2</sub>O to provide the desired almost analytically pure (elemental C, H, and N for MoC<sub>28</sub>H<sub>39</sub>F<sub>6</sub>NO<sub>8</sub>S<sub>2</sub>) key intermediate 2,6-di-iso-propylphenylimino-neophylidene-bis(triflate)-dimethoxyethane molybdenum(IV) complex [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(=CHMePh)(OTf)<sub>2</sub>DME] (5.94g, 76%, in three crops). Its <sup>1</sup>H NMR (400MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 14.45 (s, 1H, CHCMe<sub>2</sub>Ph), 7.57 (d, 2H, aromatic), 7.18 (t, 2H, aromatic), 6.97-6.89 (m, 4H, aromatic), 3.84 (sept, 2H, CHMe<sub>2</sub>), 3.73 (s, 3H, OCH<sub>3</sub>), 3.18 (m, 2H, OCH<sub>2</sub>), 2.84-2.78 (m, 5H, OCH<sub>3</sub>, OCH<sub>2</sub>), 1.91 (s, 6H, CHCMe<sub>2</sub>Ph), 1.37 (d, 6H, CHMe<sub>2</sub>), 1.21 (d, 6H, CHMe<sub>2</sub>) and the <sup>13</sup>C NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 328.4, 152.1, 151.8, 148.7, 130.6, 128.7, 126.6, 124.3, 72.8, 70.0, 65.7, 61.9, 58.8, 31.1, 28.3, 25.6, 22.8. [Schrock et al. *J Am Chem Soc* **112** 3875 1990.]

The Schrock-Hoveyda catalyst is prepared in a glove box or in Schlenk equipment under an argon or N<sub>2</sub> atmosphere by adding solid KH (360mg, 9mmol, 3 equivalents) in portions to a stirred solution of H<sub>2</sub>[S-BIPHEN] (1.06g, 3mmol, see [R- 329735-68-4, S- 205927-03-3, RS- 101203-31-0] in Part 2) in THF (100ml), whereby H<sub>2</sub> gas evolves while the K<sub>2</sub>[BIPHEN] salt is being formed. After stirring for 18 hours at ~25°, the suspension is filtered through Celite to remove excess of KH, and the key intermediate above [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(=CHMePh)(OTf)<sub>2</sub>DME] (2.24g, 2.83mmol, 0.94 equivalents) is added as a solid during 2 minutes. The mixture is stirred for 3 hours at ~25°, the volatiles are removed *in vacuo*, and the residual red powder is treated with \*C<sub>6</sub>H<sub>6</sub> (40ml), and the slurry is filtered through Celite to remove insoluble CF<sub>3</sub>SO<sub>3</sub>K and washed with \*C<sub>6</sub>H<sub>6</sub> until the washings are no longer red in colour. The red \*C<sub>6</sub>H<sub>6</sub> solutions are combined, evaporated *in vacuo*, and the spongy red solid is recrystallised from Et<sub>2</sub>O (~4ml) to give the desired catalyst, mainly in the *syn* configuration at the alkylidene group, as dark orange red crystals (1.35g, 64%). Its <sup>1</sup>H NMR (400MHz, \*C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 10.98 (s, 1H, J<sub>CH</sub> = 123Hz, *syn* alkylidene C=CHCMe<sub>2</sub>Ph), 7.42 (m, 3H, biphenyl and Ph CHs), 7.16 (m, 3H, biphenyl and Ph CHs), 7.05 (br t, J = 7.6Hz, 1H, aromatic), 6.92 (s, 3H, aromatic), 3.70 (heptet, J = 7.0Hz, 2H, *iso*-CHMe<sub>2</sub>), 2.13 (s, 3H), 2.15 (s, 3H), 1.85 (s, 3H), 1.74 (s, 3H), 1.66 (s, 3H) {these signals 2.13—1.66 ppm are from 4 biphenyl-Mes and one methyl from CMeMePh}, 1.59 (s, 9H, *tert*-CMe<sub>3</sub>), 1.54 (s, 9H, *tert*-CMe<sub>3</sub>), 1.14 (d, J = 7.0Hz, 6H, *iso*-CHMe<sub>2</sub>), 1.13 (s, 3H, CMeMePh), 0.90 (d, J = 7.0Hz, 6H, *iso*-CHMe<sub>2</sub>), and the <sup>13</sup>C{<sup>1</sup>H} NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, TMS) has δ at 277.1 (d, J<sub>CH</sub> = 123Hz, *syn* alkylidene carbon), 155.4, 154.5, 154.3, 151.3, 146.8, 140.0, 138.0, 136.5, 135.7, 132.0, 131.1, 130.9, 130.6, 129.6, 128.2, 127.9, 126.3, 123.8, 53.7, 35.95, 35.7, 34.7, 33.1, 33.0, 30.9, 30.4, 29.2, 24.6, 23.0, 20.8, 20.7, 17.2, 16.7, 14.6. [Alexander et al. *J Am Chem Soc* **120** 4041 1998.] See also a slightly modified preparation in which the red spongy powder is dissolved in Et<sub>2</sub>O (18ml), transferred to a 20ml vial, kept uncapped in a well purged glovebox (with N<sub>2</sub>) until the volume decreased to 5ml, the red solution is decanted from the red crystal blocks which were washed with cold Et<sub>2</sub>O and dried *in vacuo* to provide a 72% yield of analytically pure (elemental C, H, and N for MoC<sub>46</sub>H<sub>61</sub>NO<sub>2</sub>) *syn*-(*S*)-catalyst. The X-ray crystal structure proved the absolute configuration of the biphenolate ligand. Related complexes in which the 2,6-di-iso-propylphenylimino ligand is replaced by various other ligands have been similarly prepared. [Alexander et al. *Organometallics* **19** 3700 2000.] Schrock-Hoveyda catalysts with *R*-BIPHEN and *racemic*-BIPHEN also form red crystals, and have been similarly prepared. All are *air* and *moisture sensitive*, and have to be stored cold and in the dark. NMR data revealed that the *syn*-alkylidene conformer predominates in solution with a *syn-anti* exchange rate of ~1 s<sup>-1</sup>. However, evidence from ROM (ring-opening metathesis) polymerisation suggests that reactions with the *anti* conformer are catalytically more active than with the *syn* conformer by some orders of magnitude [Oskam & Schrock *J Am Chem Soc* **115** 11831 1993].

The related catalyst where the *tert*-buyl groups in the biphenol ligand are replaced by 1-adamantyl groups Mo(*N*-2,6-di-iso-Pr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(CHCMe<sub>2</sub>Ph)(*syn*-(*S*)-*tert*-Bu<sub>2</sub>Me<sub>4</sub>(BIAD)) has been prepared in 54% yield as orange crystals from the above key intermediate [Mo(=N-2,6-isoPr<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)(=CHMePh)(OTf)<sub>2</sub>DME] and (*S*)-H<sub>2</sub>(BIAD) (see in Part 2) except that benzylpotassium was used instead of KH to prepare the phenolate salt. X-ray crystallography of the catalyst also confirmed the *S*-absolute configuration of (-)-H<sub>2</sub>(BIAD). [Alexander et al. *Organometallics* **19** 3700 2000; and for the X-ray crystallographic structures of related Mo-2,2'-dihydroxy-1,1'-

biaryls see Totland et al. *Macromolecules* **29** 6114 1996.]

These and related chiral molybdenum complexes are remarkable catalysts that effect efficient asymmetric enantioselective ring-closing olefin metathesis (RCM) [Alexander et al. *J Am Chem Soc* **120** 4041 1998, Weatherhead et al. *Tetrahedron Lett* **41** 9553 2000], ring-opening metathesis/cross metathesis (ROM/CM) [La et al. *J Am Chem Soc* **123** 7767 2001], as well as ring-opening olefin metathesis (ROM) polymerisation [Totland et al. *Macromolecules* **29** 6114 1996] reactions. These complexes catalyse olefin metathesis reactions in which the starting olefin is a racemic mixture, and in the reaction process *kinetic resolution* occurs, whereby the products, and any unreacted starting material, are of one chirality, i.e. the asymmetric centre in the racemic mixture is converted into one enantiomer at the expense of the other enantiomer. The above Mo-(S)-[BIPHEN] complex also catalyses efficient and enantioselective *desymmetrisation* reactions, e.g. ( $\pm$ )-3-allyloxy-2,4-dimethylpenta-1,4-diene and its 1,5-dimethyl derivative (4-allyloxy-3,5-dimethyl-hepta-2,5-diene) are cyclised to 2-(prop-2'-en-2'yl)-3(R)-methyl-2,5-dihydrofuran and 2-(but-2-en-2-yl)-3(R)-methyl-2,5-dihydro-furan in the absence of solvent, 1-2 mol% catalyst, at 22° in 5 minutes in 85% and 93% yields, and 93%ee and 99%ee respectively [La et al. *J Am Chem Soc* **120** 9720 1998]. [For a review on olefin metathesis see Grubbs *Tetrahedron* **60** 7117 2004, and R.H. Grubbs (Ed.), *Handbook of Metathesis*, Vols 1-3, Wiley-VCH, 2003. ISBN 3527306161.]

**Ferric (III) acetylacetonate [Fe(acac)<sub>3</sub>, iron(III) tris(2,4-pentadionate)]** [14024-18-1] **M 353.2, m 181.3-182.3°**. When recrystallised twice from \*benzene/petroleum ether, it has **m** 181.3-182.3° corr [Finn et al. *J Chem Soc* 1256 1938]. However, when recrystallised from EtOH or Et<sub>2</sub>O it has **m** 179° [Hantzsch & Desch *Justus Liebigs Ann Chem* **323** 13 1902]. Recrystallisation from absolute EtOH also gives material with **m** 159.5° [Emmert & Jacob *Chem Ber* **67** 286 1934]. Dry it for 1 hour at 120°. [*Beilstein* **1** I 404, **1** II 836, **1** IV 3675.]

Fe(acac)<sub>3</sub> catalyses a large variety of chemical reactions such as aromatic substitution, cross-coupling, Friedel-Crafts etc [see B. Pleitker (ed), *Iron Catalysis in Organic Chemistry* Wiley-VCH, 2008, ISBN 978-3-527-31927].

**Ferrous (II) acetylacetonate [Fe(acac)<sub>2</sub>, iron(II) bis(2,4-pentadionate)]** [14024-17-0] **M 254.1, m 175°**. For catalytic properties see B. Pleitker (ed), *Iron Catalysis in Organic Chemistry* Wiley-VCH, 2008, ISBN 978-3-527-31927, and for the preparation see Chapter 5, Metal-Organic Compounds. [*Beilstein* **1** III 3122, **1** IV 3676]

**Iridium(I) bis(1,5-cyclooctadiene) tetrafluoroborate complex [bis(1,5-cyclooctadiene)iridium (I) tetrafluoroborate, Ir(COD)<sub>2</sub><sup>+</sup>.BF<sub>4</sub><sup>-</sup>]** [35138-23-9] **M 495.4, m ~190° (dec)**. This iridium complex is prepared and purified exactly as for the corresponding Rh(COD)<sub>2</sub><sup>+</sup>.BF<sub>4</sub><sup>-</sup> [35138-22-8, below] from the dimer {[Ir(COD)Cl]<sub>2</sub> [12112-67-3] below, Herde et al. *Inorg Synth* **15** 18 1974]} with COD and AgBF<sub>4</sub> in 93% yield. [Schenck et al. *Inorg Chem* **24** 2334 1985.]

**Iridium(I) chloride 1,5-cyclooctadiene complex dimer {chloro(1,5-cyclooctadiene)iridium dimer, di- $\mu$ -chlorobis[(1,2,5,6- $\eta$ )-1,5-cyclooctadiene]diiridium (I), [Ir(COD)Cl]<sub>2</sub>}** [12112-67-3] **M 671.7, m 205°(dec)**. The complex is an orange-red, air-stable solid that is soluble in \*C<sub>6</sub>H<sub>6</sub> and CHCl<sub>3</sub>, less so in Me<sub>2</sub>CO and insoluble in Et<sub>2</sub>O. It can be prepared from IrCl<sub>3</sub>.3H<sub>2</sub>O (3g) in 95% EtOH (34ml), H<sub>2</sub>O (17ml) and cycloocta-1,5-diene (6ml), through which is bubbled (with magnetic stirring) a slow stream of N<sub>2</sub> while boiling under reflux for 24 hours, during which time a brick-red product precipitates. The mixture is cooled and Ir(COD)Cl<sub>2</sub> is filtered off, washed with ice-cold MeOH to remove unreacted COD and dried *in vacuo* at 25° for 8 hour (yield 1.5g, 72%, decomp >200°). [Herde et al. *Inorg Synth* **15** 18 1974.] Alternatively, to a mixture of H<sub>2</sub>O (100ml), isoPrOH (35ml) and 1,5-cyclooctadiene (18ml) is added (NH<sub>4</sub>)<sub>2</sub>IrCl<sub>6</sub> (20g), and the mixture is refluxed under N<sub>2</sub> for 18 hours when the colour changes to orange-red, and a red or orange solid separated on cooling. This is filtered off (frit), and washed with EtOH (2 x 5ml at 0°) to give the complex in high purity (14.0g, 92%). Recrystallisation (if required) is best carried out by slowly adding an equal volume of EtOH to a saturated stirred solution of the solid in CH<sub>2</sub>Cl<sub>2</sub>, followed by gentle removal of half of the mixed solvent under reduced pressure. The stirred solution is cooled spontaneously to -30° during the process and the complex is filtered off, washed with EtOH (2 x 5ml at 0°) and dried *in vacuo*. It is identified by the characteristic IR bands (Nujol) at 907, 970, 980 and 1002 cm<sup>-1</sup>; and the vinyl CH resonance in the <sup>1</sup>H NMR (CDCl<sub>3</sub>) spectrum at  $\delta$  4.3. [Crabtree et al. *Synth React Inorg Met-Org Chem* **12** 407 1982.]

It is the metal complex precursor with allyamines for asymmetric allylic substitutions; for resting state and kinetic studies [cf Markovic & Hartwig *J Am Chem Soc* **129** 11680 2007].

**Iridium(I) (1,5-cyclooctadiene)- $\eta^5$ -(indenyl) {(1,5-cyclooctadiene)- $\eta^5$ -(indenyl)iridium (I), [(COD)(C<sub>9</sub>H<sub>7</sub>)Ir(I)], [(Ind)Ir(I) (COD)]}** [102525-11-1] **M 414.5, m 126-131°**. The complex is obtained by adding solid indenyllithium (0.73g, 5.98mmol) to a solution of the preceding dimer [Ir(COD)Cl]<sub>2</sub> (2.0g, 2.98mmol) in THF (50ml) and stirring at ~20° for 1 hour. The solvent is removed *in vacuo*, extracted with pentane (6 x 50ml) and the combined extracts are slowly evaporated to deposit pale yellow crystals (2.15g, 86%). It has <sup>1</sup>H NMR (200MHz) with  $\delta$  at 1.65-1.86 (m, 8H, COD CH<sub>2</sub>), 3.99 (m, 4H, COD CH), 5.00 (d, 2H, indenyl H1/H3), 5.76 (t, 1H, indenyl H2), 6.98-7.16 (m, 4H, indenyl H4-H7); and <sup>13</sup>C NMR (50MHz) with  $\delta$  at 34.36 (COD CH<sub>2</sub>), 51.35 (COD CH), 72.80 (indenyl C1/C3), 85.15 (indenyl C2), 110.23 (indenyl C8/C9), 121.75 and 124.89 (indenyl C4-C7). [Merola & Kacmarcik *Oganometallics* **8** 778 1989, Abad *Inorg Chim Acta* **121** 213 1986, Crabtree et al. *Synth React Inorg Met-Org Chem* **12** 407 1982, Naderer Siebel *Organomet Chem* **518** 181 1996.] [Uson & Oro *Inorg Synth* **23** 126 1985.]

[(Ind)Ir(I) (COD)] is a powerful C-H activation catalyst for preparing phenols from arenes in the presence of pinacolborane and bis(diphenylphosphino)ethane that borylate arenes to the intermediate pinacol arylboryl esters which are then converted to the respective phenols by reaction with *oxone* [2KHSO<sub>5</sub>.KHSO<sub>4</sub>. K<sub>2</sub>SO<sub>4</sub>] in Me<sub>2</sub>CO at 25° within ~7 minutes in 51 to 88% yields depending on the substituents [Maleczka et al. *J Am Chem Soc* **125** 7792 2003.]

**Iridium(I)  $\mu$ -chloro-bis(cyclooctene) dimer {di- $\mu$ -chlorotetrakis(cyclooctene)diiridium(I), [Ir(COE)<sub>2</sub>Cl]<sub>2</sub>}** [12246-51-4] **M 967.2, m 150°(dec)**. This solid yellow iridium complex is air sensitive and decomposes to a dark green solid which eventually turns black. However, it can be stored *in vacuo* in a desiccator for prolonged periods of time. It is preferable to store it in aliquots in sealed ampoules under an inert atmosphere or in a vacuum. It is soluble in Me<sub>2</sub>CO, CHCl<sub>3</sub> and \*C<sub>6</sub>H<sub>6</sub>, but it oxidises more readily in solution than in the solid state, and the necessary precautions to strictly exclude air should be exercised.

It is prepared by stirring (stirrer bar) a mixture IrCl<sub>3</sub>.3H<sub>2</sub>O (2.0g), *iso*-propanol (22ml), H<sub>2</sub>O (8ml) and cyclooctene (40ml), while a slow stream of dry N<sub>2</sub> is bubbled through the solution, and refluxed at 78° for 3 hours. During this time the colour of the solution changes from dark red to orange-yellow and the complex separates from the solution which is cooled to ~25°. It is collected rapidly by filtration, washed rapidly with ice-cold MeOH to remove excess of cyclohexene (this is done preferably in a dry box under N<sub>2</sub>). Drying *in vacuo* at 25° for 4 hours gives pure di- $\mu$ -chlorotetrakis(cyclooctene)diiridium(I) (1.5g, 59%), which decomposes at 150° and has the correct elemental analysis for C, H, Cl and Ir. [Herde et al. *Inorg Synth* **15** 18 1974, Herde & Senoff *Inorg Nucl Chem Lett* **7** 1029 1971.] By using (NH<sub>4</sub>)<sub>3</sub>IrCl<sub>6</sub> (6g) suspended in an oxygen-free suspension of *iso*-PrOH (30ml) and H<sub>2</sub>O (90ml), and cyclooctene (12ml), and refluxing with stirring for 3-4 hours under N<sub>2</sub>, followed by cooling, provides an orange oil which solidifies, is collected, washed with cold EtOH, and allowed to crystallise from EtOH, collected, dried *in vacuo*, all under N<sub>2</sub>, to give of the [Ir(COE)<sub>2</sub>Cl]<sub>2</sub> (80-92% yield) as yellow needles. It should be stored under N<sub>2</sub> at room temperature. [van der Ent & Onderdelinden *Inorg Synth* **28** 91 1989.]

Although this compound is air sensitive, with the correct stoichiometry, it forms a less sensitive stable tris-boryl complex with pinacolborane (Bpin) and 4,4'-di-*tert*-butyl-2,2'-dipyridine (dtbpy) which has been crystallised, and its X-ray analysis shows that it has the composition [Ir(dtbpy)(COE)(Bpin)<sub>3</sub>]<sup>‡</sup>. These complexes are catalytically active in the C-H borylation of arenes. Thus dissolution of the latter complex in \*C<sub>6</sub>D<sub>6</sub> (or \*C<sub>6</sub>H<sub>6</sub>) provided \*C<sub>6</sub>D<sub>5</sub>-Bpin within minutes at room temperature and in 80% yield. Kinetic studies revealed a kinetic isotope effect *k<sub>H</sub>/k<sub>D</sub>* of 3.8. The iridium complexes containing dtbpy are easily prepared and are air-stable. The reaction of bis(pinacolato)diboron (B<sub>2</sub>pin<sub>2</sub>) and ½[IrCl(COE)<sub>2</sub>]<sub>2</sub>/dtbpy (5 mol% of Ir) in 60 equivalents of \*C<sub>6</sub>H<sub>6</sub> at 25° for 4.5 hours gives Ph-Bpin in 83% yield; and with ½[IrCl(COE)<sub>2</sub>]<sub>2</sub>/dtbpy (0.02 mol% of Ir) at 100° for 16 hours gives Ph-Bpin in 80% yield with 8000 turnovers. These catalysts provide a simple and direct route for the synthesis of arylboronates—which were previously obtained by transmetalation with aryl lithium or arylmagnesium reagents and trialkylborates. [Ishiyama et al. *J Am Chem Soc* **124** 390 2002, cf Ishiyama et al. *Angew Chem In Ed* **41** 3056 2002.]

<sup>‡</sup>[Ir(dtbpy)(COE)(Bpin)<sub>3</sub>] is prepared in a dry box under N<sub>2</sub>; a glass flask containing a mixture of [Ir(COE)<sub>2</sub>Cl]<sub>2</sub> (350mg, 390mmol), dtbpy (209mg, 778mmol) and B<sub>2</sub>pin<sub>2</sub> (494mg, 1.95mmol), to which is added mesitylene (or xylene) (50ml) and heated with very slow stirring at 50° for 5 hours. The solvent is evaporated off at ~25°, the residue is dissolved in Et<sub>2</sub>O and allowed to evaporate slowly at ~25° to afford red cubes (blocks) of the complex

in 28% yield (52mg). Its  $^1\text{H}$  NMR (400MHz, cyclohexane- $d_{12}$ ) has  $\delta$  at 1.15 (s, 12H), 1.17 (s, 12H), 1.18 (s, 12H), 1.33 (m, 12H), 1.41 (s, 18H), 3.74 (d, 2H,  $J = 10.8\text{Hz}$ ), 7.09 (dd, 2H,  $J = 6.4, 2.0\text{Hz}$ ), 7.93 (s, 2H), 9.45 (d, 2H,  $J = 6.4\text{Hz}$ ) from  $\text{Me}_4\text{Si}$ ; and its  $^{11}\text{B}$  NMR (128MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 37 from external  $\text{BF}_3\text{OEt}_2$ . [Ishiyama et al. *J Am Chem Soc* **124** 390 2002.]

**Iridium(I) (methoxy)(1,5-cyclooctadiene) dimer {bis( $\eta^4$ -1,5-cyclooctadiene) di- $\mu$ -methoxy diiridium (I) dimer,  $[(\text{COD})(\text{OMe})\text{Ir}(\text{I})]_2$ } [12148-71-9] M 662.9, m 154-179° (dec). Deoxygenated solvents should be used and reactions should be carried out under  $\text{N}_2$  or argon. The methoxylated catalyst is prepared in Schlenk equipment in an inert atmosphere by adding a suspension of  $[\text{Ir}(\text{COD})\text{Cl}]_2$  (148mg, 222mmol, [12112-67-3]) in MeOH (10ml) to KOH (25mg, 445mmol) in MeOH (5ml), whereby the colour becomes orange-red and a yellow crystalline solid is formed. Stir for 30 minutes then add  $\text{H}_2\text{O}$  (40ml), collect the solid, wash it with  $\text{H}_2\text{O}$  (6 x 5ml), and dry it over  $\text{P}_2\text{O}_5$  *in vacuo* to give  $[\text{Ir}(\text{COD})(\text{OMe})]_2$  (124mg, 85%, decomposing at 145-155° becoming black at 155°) as a yellow air-stable solid. It is soluble in chlorinated solvents to give solutions which are air sensitive. It is soluble in MeOH,  $\text{Me}_2\text{CO}$ , hexane,  $^*\text{C}_6\text{H}_6$  and  $\text{Et}_2\text{O}$  but insoluble in  $\text{H}_2\text{O}$ . Its IR (Nujol) has  $\nu_{\text{max}}$  at 1325m, 1300m, 1232w, 1208m, 1172w, 1158w, 1060vs, 1005m, 972s, 913m, 824w, 827w, 811w, 783w, 574s, 560sh, 532m, 512w, 434w, 337m  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR ( $\text{CDCl}_3$ ) has  $\delta$  at 3.57 (8H, vinyl), 2.22 (8H, allylic H), 1.45 (8H, allylic H) and 3.28 (sharp singlet, 6H, MeO). [Uson et al. *Inorg Synth* **23** 127 1981.]**

Alternatively,  $[\text{Ir}(\text{COD})(\text{OMe})]_2$  can be prepared by boiling  $[\text{Ir}(\text{COD})\text{Cl}]_2$  (1g) in MeOH (40ml) under reflux with anhydrous  $\text{Na}_2\text{CO}_3$  (0.6g) for 1 hour, filtering hot, cooling, and the yellow plates of the desired complex are filtered off, washed with MeOH and dried *in vacuo*. [Chatt & Venanzi *J Chem Soc* 4735 1975.]

It has been used successfully, by preparing it *in situ*, on adding NaOMe to the solution containing  $[\text{Ir}(\text{COD})\text{Cl}]_2$  (see below).

$[\text{Ir}(\text{I})(\text{OMe})(\text{COD})]_2$ , like the preceding complex, is a powerful C-H activation catalyst for preparing phenols from arenes in the presence of pinacolborane and 3,3'-di-*tert*-butyl-1,1'-bipyridyl which borylates arenes to the intermediate pinacol arylboronate esters which are then converted to the respective phenols by reaction with *oxone* [ $2\text{KHSO}_5 \cdot \text{KHSO}_4 \cdot \text{K}_2\text{SO}_4$ ] in  $\text{Me}_2\text{CO}$  at 25° within 7 minutes in ~70% yields depending on the substituents [Maleczka et al. *J Am Chem Soc* **125** 7792 2003.] The catalyst, prepared *in situ*, in the presence of bipyridines has been used as in the previous reference for borylation of arenes in high yields and should be useful for preparing arylboronates [Ishiyama et al. *Angew Chem, Int Edn* **41** 3056 2002.]

**Iridium(III) (1,1,1,5,5,5-trifluoroacetylacetonato)(bis-cyclooctene) [(hfacac)(COE) $_2$ Ir]** [58616-58-3] M 620.5, m 96-97°. It is prepared by adding  $\text{Tl}(\text{hfacac})^\ddagger$  (600mg) to a suspension of  $[\text{Ir}(\text{COE})_2\text{Cl}]_2$  (920mg, see above) in pentane (100ml) at ~25°, and stirring for 3 hours then filtering. The red filtrate is concentrated to 10ml, kept at -30° for 24 hours and the red crystals of  $[(\text{hfacac})(\text{COE})_2\text{Ir}]$  (786.5g, 70%) are collected and dried *in vacuo* to give analytically pure complex (C and H). It is a monomeric complex in  $^*\text{C}_6\text{H}_6$  and has IR (KBr) with  $\nu_{\text{max}}$  at 2960m, 2910s, 2840s, 2660w, 1620m, 1580s, 1540s, 1500w, 1440s, 1340s, 1310w, 1250s, 1210w, 1140s, 1040w, 1015w, 970w, 950m, 945w, 910s, 895w, 875w, 860w, 850w, 810m, 795s, 740m, 680s, 630m, 590s, 570m, 545s, 435m  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 6.26 (s, 1H), 2.8-2.4 (br m, 4H), 2.3-1.8 (br m, 8H), 1.7-1.3 (br m, 16H) from TMS.

$(\text{Acac})(\text{COE})_2\text{Ir}$  is similarly prepared from  $\text{Tl}(\text{acac})^\ddagger$  {[25955-51-6]. M 303.5, m 134° obtained as for  $\text{Tl}(\text{hfacac})$  below} and  $[\text{Ir}(\text{COE})_2\text{Cl}]_2$  in 86% yield forming orange-yellow crystals with m 114-115°, M 512.6, and has IR (KBr) with  $\nu_{\text{max}}$  at 2970s, 2910s, 2890s, 2830s, 1620m, 1570s, 1540m, 1510s, 1460s, 1440m, 1410w, 1370m, 1340w, 1310w, 1270w, 1185w, 1140w, 1015s, 970w, 930w, 910w  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 5.43 (s, 1H), 1.92 (s, 6H), 2.4-1.7 (br m, 8H), 1.7-1.0 (br m, 20H) from TMS. [Diversi et al. *J Organomet Chem* **125** 253 1977.]

$^\ddagger$ Thallium 1,1,1,5,5,5-hexafluoroacetylacetonate  $[\text{Tl}(\text{hfacac})]$  is obtained when to a solution of thallos ethoxide (8g, prepared as in Fieser & Fieser's *Reagents for Organic Synthesis* **2** 407 1969, see [20398-06-5] in Chapter 5, Metal-Organic Compounds] in EtOH (60ml) is added hexafluoroacetylacetone (6.7g, [see 1522-22-1]) at ~25°, and after 30 minutes the solvent is removed *in vacuo* and the residue is washed with pentane and dried to give  $\text{Tl}(\text{hfacac})$  (11.8g, 95%) as a microcrystalline solid. [Ingrosso et al. *J Organomet Chem* **84** 75 1975.]

**Iron(0) 2,2'-bipyridine [bpyFe(0)]** [*bpyFe(1<sup>+</sup>)* 500295-31-8; *bpyFe(2<sup>+</sup>)* 73871-24-6; *bpyFe(3<sup>+</sup>)* 51232-88-3] **M 212.0**. This iron catalyst mediates *ene-carbocyclisations* such as formal [4+4] ene reaction of trienes [Takacs & Anderson *J Am Chem Soc* **109** 2200 1987, Takacs et al. *Tetrahedron* **46** 5507 1990], regio- and chemo-selective diene to olefin cross coupling reactions [Takacs et al. *Organometallics* **5** 2395 1986], stereoselective and regiocontrolled formation of substituted tetrahydropyrans [Takacs et al. *Tetrahedron Lett* **28** 5627 1987]. The “active bpyFe(0) catalyst” is prepared *in situ* from Fe(acac)<sub>2</sub> (see [14024-17-0]) generated by a minimum of 3.0 equivalents of Et<sub>3</sub>Al and Fe(acac)<sub>3</sub> (see [14024-18-1], in \*C<sub>6</sub>H<sub>6</sub> at 0°), and 2,2'-bipyridine (1:1 with respect to Fe, sublimed at 65°/0.01mm) to which is added the olefin(s) at or near room temperature. [The Fe(acac)<sub>3</sub> is recrystallised from EtOH or \*C<sub>6</sub>H<sub>6</sub>/hexanes, dried at 25°/0.01mm, or sublimed at 100°/0.05mm.] Prior to the reduction of the iron, 1.1 equivalents of 2,2'-bipyridine and 10 equivalents of “addend” (furan or 2-methylfuran) are added to the Fe(acac)<sub>3</sub> in \*C<sub>6</sub>H<sub>6</sub>. After reduction, the substrate(s) are added and the reaction is allowed to proceed at 25° and monitored. The “addend” is added mainly to slow down the disproportionation of the active bpyFe(0) into inactive bpy<sub>2</sub>Fe(0) and Fe<sub>metal</sub> [Takacs et al. *Tetrahedron* **46** 5507 1990]. **Note:** The catalytically active 1:1 complex bpyFe(0) has not been isolated, but is assumed to be in solution and is most probably stabilised in solution by coordinating with some entity such as ethylene (derived from Et<sub>3</sub>Al), or perhaps even furan or methylfuran as these are required for its stability (see below).

The reduction of FeCl<sub>3</sub> by 3.1 equivalents of *iso*-PrMgBr in the presence of 1.1 equivalents of 2,2'-bipyridine and 2,3-dimethylbutadiene yields an active catalyst by forming the **bpyFe(0)2,3-butadiene complex** [104714-94-5] in which the olefin may be involved in the reaction [Takacs et al. *Organometallics* **5** 2395 1986].

Carbocyclisations are also mediated by Pd, Ni, Co, and Rh complexes [for bibliography see Takacs et al. *Tetrahedron* **46** 5507 1990]. **Note:** the Fe(0) valence is dubious.

**(2,3-O-Isopropylidene)-2,3-dihydroxy-1,4-bis-(diphenylphosphino)butane (DIOP)** [*4R,5S(-)-* 32305-98-9, *4S,5R(+)-* 37002-48-5] **M 498.5, m 88-89°, 88-90°, [α]<sub>D</sub><sup>19</sup> (-) and (+) 26° (c 2.3, CHCl<sub>3</sub>), also 4R,5S- has [α]<sub>D</sub><sup>22</sup> (-) 12.3° (c 4.6, \*C<sub>6</sub>H<sub>6</sub>), pK<sub>Est</sub> ~0.0**. These are quite stable in air and have been recrystallised from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. After 2 recrystallisations from EtOH, they are generally pure by TLC on silica gel using Me<sub>2</sub>CO/hexane as solvent. The optical purity can be determined by GC [Chiral Select 1000 column 15m x 0.25mm using He carrier gas (1ml/minute); or Chirasil-L-Val fused silica 25m x 0.25mm] or by HPLC using an (S,S)-Whelk-01 column (5.0μm, 25cm x 0.46cm). [Details of the general preparation of DIOP are described under **3,4-O-isopropylidene-(3S,4S)-dihydroxy-(2R,5R)-bis(diphenylphosphino)hexane (RSSR-dimeDIOP)**, see [258873-45-9]] in Part 2. The rhodium(I) complexes described below catalyse the asymmetric hydrogenation of β-substituted α-acetamidoacrylic acids to the corresponding chiral α-amino acid derivatives with high stereoselectivity. [Kagan & Dang *J Am Chem Soc* **94** 6429 1972.]

The DIOP derived catalysts used are **Rh(I)<sup>+</sup>(DIOP)(COD) X<sup>-</sup>** where X<sup>-</sup> is Cl<sup>-</sup>, BF<sub>4</sub><sup>-</sup>, SbF<sub>6</sub><sup>-</sup> or PF<sub>6</sub><sup>-</sup> and are prepared *in situ* by reaction of Rh(I)<sup>+</sup>(COD)<sub>2</sub> X<sup>-</sup> with the DIOP ligand in a solvent (CH<sub>2</sub>Cl<sub>2</sub>, MeOH, PhCH<sub>3</sub>, THF or \*C<sub>6</sub>H<sub>6</sub>), using a DIOP/Rh ratio of <2 (preferably 1.1:1.0) to avoid the formation of Rh(I)<sup>+</sup>(DIOP)<sub>2</sub> which exhibits very little, if any, catalytic activity due to its inability to coordinate with the olefinic substrate. A **typical hydrogenation** is carried out in a Fisher-Porter tube which is filled (in a dry-box) with the enamide substrate (0.25mmol), the appropriate solvent (5ml), and pre-formed Rh(I)<sup>+</sup>(DIOP)(COD) X<sup>-</sup> [Rh precursor, Rh(I)<sup>+</sup>(COD)<sub>2</sub> X<sup>-</sup> (0.005mol) in the appropriate solvent (3ml), DIOP (0.055ml of 0.1M solution in, e.g. PhCH<sub>3</sub>, 0.0055mmol) and stir for 10 minutes]. After sealing, the tube is removed from the dry-box, placed behind a proper body shield, and *ca* five vacuum/refilling cycles of H<sub>2</sub> are performed. Hydrogenation is performed at room temperature, with the desired pressure of H<sub>2</sub> (1.1 to 50 bar, usually ~30psi), and is stirred vigorously for ~10 hours (but could be up to 60 hours). The reaction is completed by releasing the H<sub>2</sub>, the catalyst is filtered off through a short silica column, and the filtrate is worked up in the appropriate manner. [Kagan & Dang *J Am Chem Soc* **94** 6429 1972, Li & Zhang *J Org Chem* **65** 5871 2000, Yan & RajanBabu *Org Lett* **26** 4137 2000.] A variety of DIOP derivatives have been synthesised and tested with pre-Rh catalyst salts having different X<sup>-</sup> counter ions in several solvents. The combination which gave high yields and ee excesses of ~98% was the all-equatorial **3,4-O-isopropylidene-(3S,4S)-dihydroxy-(2R,5R)-bis(diphenylphosphino)hexane**, i.e. **RSSR-dimeDIOP**, X<sup>-</sup> = BF<sub>4</sub><sup>-</sup> and Rh(I)<sup>+</sup>(COD)<sub>2</sub> BF<sub>4</sub><sup>-</sup> [35138-22-8], in CH<sub>2</sub>Cl<sub>2</sub>, at 20-40psi and 10 hours [Yan & RajanBabu *Org Lett* **26** 4137 2000].

Other examples of the application of Rh-catalysed reactions requiring the DIOP ligand include regioselective hydroformylation of methyl *N*-aminoacrylate [and involving RhH(CO)(DOP)] [Gladiali & Pinna *Tetrahedron*:

*Asymmetry* **1** 693 1990], 1,4-disilylation of  $\alpha$ ,  $\beta$ -unsaturated ketones [Matsumoto & Hayashi *Tetrahedron* **50** 335 1994], the Heck reaction with alkenyl iodides [Sato et al. *Tetrahedron* **50** 371 1994], as well as the asymmetric hydrogenation of enamides [Morimoto et al. *Tetrahedron: Asymmetry* **6** 23 1995, Morimoto et al. *Tetrahedron Asymmetry* **6** 75 1995].

**Lanthanide trifluoromethanesulfonate (triflate) salts.** These behave mostly as water-tolerant *Lewis acids* which catalyse a variety of organic functional group transformations. The lanthanide triflates below have been prepared and purified by the general procedure described here unless otherwise stated. The salts are prepared by adding excess of lanthanide(III) or (IV) oxide (>99% purity, 30mmol) to an aqueous solution of trifluoromethanesulfonic acid (50% v/v, 21.2ml) and heating to boiling for 30 minutes to 1 hour (*alternatively*, at 100° for 2 hours). The mixture is filtered to remove the unreacted oxide. The water is then removed under a vacuum. The resulting hydrated salt is dried by heating under vacuum at 180 to 200° for 48 hours (200°/0.5mm for 40 hours was also reported). These salts are quite hygroscopic and all manipulations of reactions using these salts should be carried out using Schlenk equipment or glove boxes under an inert atmosphere (N<sub>2</sub>, He or Ar) to prevent contamination with H<sub>2</sub>O which will lead to the formation of insoluble lanthanide hydroxides, and to maximise the activity of the catalyst. [Forsberg et al. *J Org Chem* **52** 1017 1987, Kobayashi & Hachiya *J Org Chem* **59** 3590 1994.]

**Cerium(IV) trifluoromethanesulfonate [cerium triflate, Ce(OSO<sub>2</sub>CF<sub>3</sub>)<sub>4</sub> H<sub>2</sub>O, Ce(OTf)<sub>4</sub> H<sub>2</sub>O] [698999-65-4] M 736.4 (anhydrous), slowly decomposes above 120° to give trivalent cerium species.** Preparation of Ce(OTf)<sub>4</sub> from CeO<sub>4</sub> or Ce(OH)<sub>4</sub> and trifluoromethanesulfonic acid as above for lanthanide triflates is usually unsuccessful. It is best prepared by adding, with vigorous stirring, a solution of K<sub>2</sub>CO<sub>3</sub> (17.3g, 120mmol) in H<sub>2</sub>O (95ml) to a solution of cerium(IV) ammonium nitrate (27.4g, 50mmol) [16774-21-3] in H<sub>2</sub>O (80ml), whereupon a pale yellow carbonate separates. This is filtered off and washed several times with H<sub>2</sub>O. CF<sub>3</sub>SO<sub>3</sub>H (triflic acid, 17.7ml, 200mmol) is slowly added at 0° to this “wet” cerium(IV) carbonate whereby it dissolves, and the resulting orange coloured solution is evaporated under reduced pressure, and the residue is dried *in vacuo* at 70° for 10 hours to give Ce(OTf)<sub>4</sub> as a yellow powder (36.2g). [Note that a “wet” carbonate is important as very dry carbonate does not react readily with triflic acid at 0°.] The IR (KBr) has characteristic bands at  $\nu_{S=O}$  1230 and  $\nu_{C-F}$  1010 cm<sup>-1</sup> and it analyses as a *monohydrate* (by Karl Fischer). The salt is hygroscopic and should be stored under N<sub>2</sub> and preferably in aliquots in sealed containers. It is soluble in H<sub>2</sub>O, EtOH, THF, 1,2-dimethoxyethane and dioxane, but almost insoluble in hexane, \*C<sub>6</sub>H<sub>6</sub> and CHCl<sub>3</sub>. It has very good oxidising ability, thus converting benzylic alcohols to aldehydes or ketones, and benzylic type CH<sub>2</sub> to CO, e.g. substituted toluenes to their corresponding benzaldehydes, ethylbenzene to acetophenone and diphenylmethane to benzophenone in high yields. [Imamoto et al. *Chem Lett* 1445 1990]. Ce(OTf)<sub>4</sub> is an efficient catalyst for the ring opening of epoxides with high regio and stereo selectivity. Ring opening, e.g. of styrene oxides (phenyloxiranes), yields essentially *trans* products with the OH, OR or OAc entering groups (when H<sub>2</sub>O, alcohols or AcOH are used) attacking the ‘benzylic carbon’ atom and with high optical purity when the reaction is carried out at ~-10° [e.g. *R*(+)-styrene oxide provides *S*(+)-PhCH(OMe)CH<sub>2</sub>OH]. Ring opening of thiiranes with H<sub>2</sub>O, alcohols or AcOH also yield *trans* products which dimerise to the corresponding dithianes. [Iranpoor et al. *Synth Commun* **28** 347 1998, cf Vougiokas & Kagan *Tetrahedron Lett* **28** 6065 1987.]

**Gadolinium(III) trifluoromethanesulfonate [Gd(OTf)<sub>3</sub>] [52093-29-5] M 604.5, decomposes >120°.** Gd(OTf)<sub>3</sub> is prepared by the general method described above. If suspect, then add aqueous triflic acid (50% v/v) and proceed as above. It catalyses the aminolysis of epoxides in an extraordinarily efficient manner in aprotic solvents (e.g. toluene, CH<sub>2</sub>Cl<sub>2</sub>) with complete *trans* stereoselectivity and high regioselectivity [Chini et al. *Tetrahedron Lett* **35** 433 1994]. It also catalyses the reactions between nitriles and amines to yield a variety of amidines which, depending on the amine, can be used to prepare cyclic amidines, pyrimidines and *s*-triazines [Forsberg et al. *J Org Chem* **52** 1017 1987]. It is a water-tolerant *Lewis acid* used in aldol reactions of silyl enol-ethers and aldehydes in ~79-89% yields (see below) [Kobayashi & Hachiya *J Org Chem* **59** 3590 1994].

**Hafnium(IV) trifluoromethanesulfonate hydrate [Hf(OTf)<sub>4</sub>] [161337-67-3] M 774.8 (anhydrous),**

**m >350°.** Hf(OTf)<sub>4</sub> is prepared by the general method described above. If suspect, then add aqueous triflic acid (50% v/v) and proceed as in the general method above. It is an efficient catalyst in the Fries rearrangement of acyloxy benzene or naphthalene derivatives, and for the regioselective direct acylation of phenol and naphthol derivatives with acid chlorides [Kobayashi et al. *Tetrahedron Lett* **37** 2053 1996]. It is an excellent and recyclable catalyst for mono-nitration of *o*-nitrotoluene [Waller et al. *Tetrahedron Lett* **39** 1641 1998].

**Neodymium(III) trifluoromethanesulfonate [Nd(OTf)<sub>3</sub>] [34622-08-7] M 698.3, decomposes >120°.** Nd(OTf)<sub>3</sub> is prepared by the general method described above. If suspect, then add aqueous triflic acid (50% v/v) and proceed as above. It catalyses the aminolysis of epoxides in an extraordinarily efficient manner in aprotic solvents (e.g. toluene, CH<sub>2</sub>Cl<sub>2</sub>) with complete *trans* stereoselectivity and high regioselectivity [Chini et al. *Tetrahedron Lett* **35** 433 1994]. It also catalyses the reactions between nitriles and amines to yield a variety of amidines, which, depending on the amine, can be used to prepare cyclic amidines, pyrimidines and *s*-triazines [Forsberg et al. *J Org Chem* **52** 1017 1987]. It is a water-tolerant *Lewis acid* used in aldol reactions of silyl enol-ethers and aldehydes in ~83-85% yields (see below) [Kobayashi & Hachiya *J Org Chem* **59** 3590 1994].

**Ytterbium(III) trifluoromethanesulfonate hydrate [Yb(OTf)<sub>3</sub> · x H<sub>2</sub>O] [252976-51-5] M 620.3 (anhydr), decomposes >120°.** Yb(OTf)<sub>3</sub> is prepared by the general method described above. If suspect, then add aqueous triflic acid (50% v/v) and proceed as above. It can be recrystallised from MeCN/CH<sub>2</sub>Cl<sub>2</sub>. It has IR (KBr) bands at 3650, 3350, 2300, 1650, 1300, 1040 cm<sup>-1</sup>; and <sup>13</sup>C NMR (270MHz, D<sub>2</sub>O) at δ 122.4 (q, *J* = 317Hz) using sodium 2,2-dimethyl-2-silapentane-5-sulfonate as internal standard.

It catalyses the aminolysis of epoxides in an extraordinarily efficient manner in aprotic solvents (e.g. toluene, CH<sub>2</sub>Cl<sub>2</sub>) with complete *trans* stereoselectivity and high regioselectivity [Chini et al. *Tetrahedron Lett* **35** 433 1994]. It also catalyses the *trans* addition of indole (at position 3) to epoxides (e.g. to phenoxymethyloxirane) in >50% yields at 60° (42 hours) under pressure (10 Kbar) and was successfully applied for an enantioselective synthesis of (+)-diolmycin A2 [Kotsuki *Tetrahedron Lett* **37** 3727 1996]. Of the ten lanthanide triflates, Yb(OTf)<sub>3</sub> gave the highest yields (> 90%, see above) of condensation products by catalytically activating formaldehyde, and a variety of aldehydes, in hydroformylations and aldol reactions, respectively, with trimethylsilyl enol-ethers in THF at room temperature. All the lanthanide triflates can be recovered from these reactions for re-use. [Kobayashi & Hachiya *J Org Chem* **59** 3590 1994.]

**Methyltrioxorhenium (MTO, trimethylrhenium(VII) trioxide, Me<sub>3</sub>ReO<sub>3</sub>) [70197-13-6] M 249.2, m 110°, 111°, pK<sub>8</sub><sup>25</sup> 7.53.** MTO is an air-stable carbon-rhenium oxide which is prepared from tetramethylrhenium oxide (Me<sub>4</sub>ReO, [53022-70-1]) or trimethylrhenium dioxide (Me<sub>3</sub>ReO<sub>2</sub>, [56090-01-8]) (100mg) in a 1L evacuated bulb, and dry air is admitted up to a pressure of ~760mm (atmospheric). After a few days long needles of Me<sub>3</sub>ReO<sub>3</sub> are formed, and after 4 weeks in excess of 50% yields of the trioxide are obtained. These can be resublimed to analytical purity *in vacuo* (at ~25°/1mm or 65°/0.001mm), have a sharp melting point (110°), are not decomposed in the gas phase <300°. **Alternatively on a larger scale**, silver tetraoxorhenate (14.0g, 39mmol, AgReO<sub>4</sub> [7784-00-1]) is dissolved in MeCN (150ml) followed by addition of Me<sub>3</sub>SiCl (10.8ml, 85mmol, TMSCl [75-77-4]) whereby a white precipitate of AgCl separates and the solution becomes orange in colour. The suspension is treated with Me<sub>4</sub>Sn (6.0ml, 43mmol, [594-27-4]), stirred for 24 hours, filtered into a Schlenk sublimation tube (~0.75L) and evaporated in a vacuum (20mbar), using a liquid N<sub>2</sub> trap (for volatile Sn compounds). A cold finger is then inserted into the tube and the Me<sub>3</sub>SnCl [1066-45-1] is sublimed out at ~25°/20mbar; this compound is removed and the vacuum is increased to 0.01mbar when MTO sublimes out as a yellow-white crystalline solid. Traces of toxic Sn compounds may be present in this MTO (causing a pungent odour), but are removed by spreading it on filter paper in a fume hood for 1-3 hours, and the purity is checked by elemental analysis. MTO (7.8g) is thus obtained in 80% yield. Similar results are obtained on a 100g scale. The recovered Me<sub>3</sub>SnCl can be converted to SnMe<sub>4</sub> by methylating with MeMgCl (23% in THF). High-purity MTO can be obtained by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/hexane (73%). The white solid turns gray on standing without alteration in its activity, but can be minimised by storing it away from light under N<sub>2</sub>. The Sn compounds are **HIGHLY TOXIC, and all work should be carried out in an efficient fume cupboard**. [Herrmann & Kratzer *Inorg Synth* **33** 111 2002, Herrmann et al. *Angew Chem, Int Ed. Engl* **36** 2652 1997.]

MTO forms colourless needles which develop a greyish tint on prolonged storage in light. It differs from methylperrhenate in being stable to hydrolysis by H<sub>2</sub>O. It is soluble in MeCN, \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub>, EtOH and H<sub>2</sub>O, but sparingly soluble in CS<sub>2</sub> and hexane. Its MS has parent ion peaks of MeReO<sub>3</sub><sup>+</sup> at 248 and 250 mass units for

the species containing  $^{185}\text{Re}$  and  $^{187}\text{Re}$  in a ratio consistent with the relative isotopic abundances. The UV/VIS ( $\text{Et}_2\text{O}$ ) has  $\lambda_{\text{max}}$  ( $\epsilon$ ) at 260nm (1400,  $\text{Lmol}^{-1}\text{cm}^{-1}$ ) and 232nm (1900,  $\text{Lmol}^{-1}\text{cm}^{-1}$ ).  $^1\text{H}$ ,  $^{13}\text{C}$  and  $^{17}\text{O}$  NMR ( $\text{CDCl}_3$ ,  $25^\circ$ ) spectroscopy shows signals with  $\delta_{\text{H}}$  at 2.61 (sharp),  $\delta_{\text{C}}$  at 19.03 ( $^2J = 138\text{Hz}$ , C-H) and  $\delta_{\text{O}}$  at 829 respectively; and the IR (hexachlorobutadiene mull) has  $\nu_{\text{max}}$  at 1360ms (*CH deformation*), 2895 and 2980 (*CH stretching*)  $\text{cm}^{-1}$ ; the IR in gaseous state (at  $70^\circ$ ) has  $\nu_{\text{max}}$  at 1003w(*ReO<sub>3</sub> sym str*), 985w, 975vs, 962vs(*ReO<sub>3</sub> antisym str*), 743mw(*CH<sub>3</sub>-rock*), 574w(*Re-C*), 324w(*ReO<sub>3</sub> deformation*)  $\text{cm}^{-1}$ , and the IR (argon matrix) 1000ms(*ReO<sub>3</sub> sym str*), 970vs and 966m(*ReO<sub>3</sub> antisym str*), 566w(*Re-C*)  $\text{cm}^{-1}$ . Its Raman spectrum in  $\text{CS}_2$  has  $\nu_{\text{max}}$  at 999s(*ReO<sub>3</sub> sym str*)  $\text{cm}^{-1}$ , and in the solid state it has  $\nu_{\text{max}}$  at 999s(*ReO<sub>3</sub> sym str*), 964m(*ReO<sub>3</sub> antisym str*), 530m(*Re-C*), 330m and 242m(*ReO<sub>3</sub> deformation*)  $\text{cm}^{-1}$ . [Beattie & Jones *Inorg Chem* **18** 2318 1979.]

**MTO** is a *catalytic oxygen carrier* in conjunction with an oxidant, e.g.  $\text{H}_2\text{O}_2$ , UHP (urea- $\text{H}_2\text{O}_2$  complex). For example 3.7mmol of 3,4,6-tri-*O*-acetyl-D-glucal in MeOH is oxidised by 7.4mmol of  $\text{H}_2\text{O}_2$  in the presence of a catalytic amount of MTO ( $3.6 \times 10^{-2}$ mmol) at  $\sim 25^\circ$  in 2 hours to give the corresponding  $\beta$ -D-*gluco*- and  $\alpha$ -D-*manno*- pyranosides (2:1) in 78% isolated yields *via* the intermediate epoxide. [Boyd et al. *Green Chem* **5** 679 2003.] **MTO** is a versatile oxidation catalyst and very effective in olefin epoxidation, olefin isomerisation, olefin metathesis, Baeyer-Villiger oxidation of ketones to lactones, aromatic oxidation to *p*-quinones, and in Diels-Alder reactions [Summarised in Herrmann *J Organomet Chem* **500** 149 1995]. It is a catalytic oxidant for the conversion of imines to nitrones [Soldaini et al. *Org Lett* **9** 473 2007, Baldwin & Long *Org Lett* **6** 1653 2004], and catalyses the efficient and stereospecific desulfurisation of thiirane (episulfides) by  $\text{Ph}_3\text{P}$  at  $\sim 25^\circ$ , and more so if MTO is pre-treated with  $\text{H}_2\text{S}$ , probably due to the formation of Re(V)S species as the active catalyst [Jacob & Espenson *JCS, Chem Commun* 1003 1999].

**Molybdenum dioxo bis(2,4-dipentandionato- $\kappa\text{O}_2,\kappa\text{O}_4$ ) [bis(acetylacetonato)dioxomolybdenum (IV), dioxobis(2,4-pentadionato)molybdenum (IV),  $\text{MoO}_2(\text{acac})_2$  [17524-05-9] M 326.2, m 179 $^\circ$ , 184-185 $^\circ$ , 184 $^\circ$  (dec).** This complex is readily prepared by dissolving powdered ammonium *paramolybdate*  $\{(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ , 3.0g, 2.4mmol, see [12054-85-2]\} in 15% aqueous  $\text{NH}_3$  (6.0ml, or 6.0ml of 58%) with stirring. The clear solution (but if some solid remains undissolved, it will dissolve when nitric acid is finally added) is treated with 2,4-pentanedione (7.0ml, 6.8mmol) with stirring and the colour turns yellow. Concentrated  $\text{HNO}_3$  (5.0ml, d 1.42) is slowly added with stirring to the mixture which warms up and the colour turns greenish yellow. The yellow complex separates as the mixture is allowed to cool to room temperature. After  $\sim 0.5$  hours the solid is filtered off washed well with  $\text{H}_2\text{O}$ , then with EtOH, and dried *in vacuo* to give a 72% yield (4.0g) of the complex which can be increased to 93% if the hot solution is kept in a freezer at  $-10^\circ$  in 18 hours. It can be recrystallised by dissolving  $\sim 2.0$ g in hot 2,4-pentanedione (6.0ml) at  $\sim 90^\circ$  then keeping at  $-10^\circ$  in 18 hours. The yellow crystals are filtered off with suction, washed with EtOH and dried *in vacuo* to give 1.8g of product. Note that if all the 2,4-pentanedione is not removed the crystals slowly develop a reddish tint. It decomposes slowly when stored in air or under fluorescent light, and turns blue. However, it is stable when stored *in vacuo* or under  $\text{N}_2$ . It is insoluble in  $\text{H}_2\text{O}$ , but slightly soluble in EtOH,  $\text{CHCl}_3$ ,  $\text{CH}_2\text{Cl}_2$  and soluble in MeCN. [Chakravarti & Bandyopadhyay *Inorg Synth* **29** 130 1992, and from  $\text{MoO}_3$  and 2,4-pentanedione see Fernelius et al. *Inorg Synth* **6** 146 1960.] Its IR (KBr) has  $\nu_{\text{max}}$  at 1560, 1500, 1350, 1260, 1015, 930 (Mo=O *symm*), 900 (Mo=O *antisymm*), 795, 665, 570 and 445  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR ( $\text{CD}_2\text{Cl}_2$ , TMS) has  $\delta$  at 2.12, 2.14 and 5.83; and the  $^{13}\text{C}$  NMR ( $\text{CD}_2\text{Cl}_2$ , TMS) has  $\delta$  at 196.7, 185.2, 104.9, 28.0 and 25.7.

**$\text{MoO}_2(\text{acac})_2$**  catalyses: (a) the oxidation of primary (e.g. cinnamyl alcohol) or secondary alcohols (e.g. 1-phenylpropan-1-ol) to their respective aldehydes or ketones in high yields using sodium percarbonate (see [15630-89-4]) as oxidant and Adogen 464 (0.2 equivalents, as phase transfer catalyst see [72749-59-8]) in boiling (6-24 hours)  $\text{CH}_2\text{Cl}_2$  [Maignein et al. *Synlett* 439 1996]; (b) the mild and efficient deprotection of acetals (e.g. benzaldehyde dimethyl acetal, dodecanal dimethyl acetal) in  $> 70$ -95% yields using  $\sim 10\%$  of catalyst in MeCN under  $\text{N}_2$  for  $\sim 4$  hours [Kantam et al. *Synth Commun* **25** 2529 1995], (c) the oxidation of 3- $\beta$ -cholesteryl esters to the corresponding 5,6- $\beta$ -epoxides in the presence of isobutyraldehyde and  $\text{O}_2$  in  $\text{CH}_2\text{Cl}_2$  in yields  $> 75\%$  [Kantam et al. *Synth Commun* **24** 961 1994], and (d) the methoxymethylation of primary, secondary and tertiary alcohols (R-OH) to their corresponding methoxymethyl ethers (R-OCH<sub>2</sub>OCH<sub>3</sub>) in 75-95% yields using excess of methylal  $\{\text{CH}_2(\text{OMe})_2$  see [109-87-5]\} in refluxing  $\text{CHCl}_3$  [Kantam & Ssnthi *Synlett* 429 1993].

**(2S)-3-exo-(Morpholino)isborneol [(-)MIB] [287105-48-0] and (2R)-3-exo-(morpholino)iso-borneol [(+)MIB] (Nugent's reagent) M 239.4, m 60-65 $^\circ$ , 65-67 $^\circ$ ,  $[\alpha]_{\text{D}}^{25} \pm 6^\circ$  (c 1, MeOH).** The (2R)-enantiomer is prepared from (2R)(+)-*cis*-3-exo-(amino)isborneol (4.53g, 26.8mmol; Chittenden & Cooper *J Chem Soc C* 49

1970) in DMSO (25ml) and  $\text{Et}_3\text{N}$  (10ml), to which is added dropwise, with stirring, a solution of di(2-bromoethyl) ether (8.07g, ~90% pure, 31.3mmol) in DMSO (20ml). After 72 hours the mixture is poured into  $\text{H}_2\text{O}$  (250ml), basified with aqueous M NaOH (60ml), extracted with  $\text{Et}_2\text{O}$  (3 x 100ml), and the extract is evaporated *in vacuo*. The residue is then dissolved in  $\text{Et}_2\text{O}$ , extracted into aqueous M HCl (50ml), basified with aqueous NaOH and re-extracted into  $\text{Et}_2\text{O}$  which is dried ( $\text{MgSO}_4$ ), filtered, evaporated *in vacuo*, and the residue is dissolved in hot hexanes (4ml/g crude), filtered and cooled to  $-30^\circ$  to give analytically pure (C, H and N) (+)MIB (2.95g, 46%). It is a white crystalline solid, **m 65-67°**, that can be stored at ambient temperature for 3 months with no alteration in spectroscopic properties, or catalytic performance when compared with freshly prepared material. It has  $^1\text{H}$  NMR ( $^*\text{C}_6\text{D}_6$ ) with  $\delta_{\text{H}}$  at 0.69 (s, 3H), 0.73 (m, 1H), 0.88 (m, 1H), 1.04 (s, 3H), 1.31 (td, 1H), 1.52 (m, 1H), 1.67 (d, 1H), 1.99 (d, 1H), 2.13 (br, 2H), 2.31 (br, 2H), 3.32-3.42 (m, 5H total), 3.92 (br d, 1H); and the  $^{13}\text{C}$  NMR ( $^*\text{C}_6\text{D}_6$ ) has  $\delta_{\text{C}}$  at 11.88, 21.10, 22.19, 27.99, 32.56, 45.35, 46.64, 49.52, 66.82, 73.37, 79.03 (and one resonance is too broad to be observed at  $\sim 25^\circ$ ). [Nugent *J Chem Soc. Chem Commun* 1369 1999.]

MIB acts as an efficient catalyst for the enantioselective addition of the ethoxy vinyl zinc reagent to aldehydes providing hydroxy vinyl ethers, which in turn, are easily converted to chiral hydroxy aldehydes [Jeon et al. *Org Lett* 7 1729 2005]. It generally catalyses the addition of organo-zinc reagents (e.g. from transmetalation of allylic boranes with alkylZn compounds) to aldehydes with very high enantiomeric preference, *viz* (*R*)-MIB providing the (*R*)-alcohol from the corresponding aldehyde [Nugent *J Chem Soc. Chem Commun* 1369 1999, Jeon et al. *Org Lett* 7 1729 2005], and this catalytic method has been adopted for the preparation of  $\alpha$ -amino acids [Chen et al. *J Am Chem Soc* 124 12225 2002], of  $\gamma$ -unsubstituted  $\beta$ -amino acids [Lurain & Walsh *J Am Chem Soc* 125 10683 2003] and for epoxy-alcohols with up to three contiguous stereocentres [Lurain et al. *J Am Chem Soc* 126 10683 2004, Lurain et al. *J Org Chem* 70 1262 2005].

**Nickel(II) acetylacetonate**  $[\text{Ni}(\text{acac})_2]$  [3264-82-2] **M 256.9, m 229-230°, b 220-235°/11mm, d<sup>17</sup> 1.455**. It is obtained by adding a solution of acetylacetone (50g, 0.5mole) in MeOH (100ml) to  $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$  (59.4g, 0.25mole) in  $\text{H}_2\text{O}$  (250ml) with stirring, followed by a solution of  $\text{NaOAc} \cdot 3\text{H}_2\text{O}$  (68g, 0.5mole) in  $\text{H}_2\text{O}$  (150ml), heating briefly on a hot plate, cooling in a refrigerator for several hours, and filtering the crystals off. Wash the emerald green solid with  $\text{H}_2\text{O}$ , dry it in a vacuum desiccator and recrystallise it from MeOH. [Charles & Pawlikowski *J Phys Chem* 62 440 1958.] The complex can be conveniently dehydrated by azeotropic distillation with toluene, and the crystals can be isolated by concentrating the toluene solution. [Wilkinson et al. *J Am Chem Soc* 76 1970 1954, *Beilstein* 1 IV 3677.] It is soluble in organic solvents such as EtOH,  $\text{CHCl}_3$ , and  $^*\text{C}_6\text{H}_6$  but insoluble in  $\text{Et}_2\text{O}$  and hexanes. Its UV spectrum has  $\lambda_{\text{max}}$  nm(log $\epsilon$ ) at 265 (4.44) and 298 (4.34) ( $10^{-4}$  M in  $\text{CHCl}_3$ ). It is a *trimer* in the solid state and a *monomer* in the vapour phase. When the metal in  $\text{Ni}(\text{acac})_2$  is coordinated with the carbenes derived from IPr.Cl, IPr. $\text{BF}_4$ , (see below) or related 1,3-dimesityl-imidazolium chloride [141556-45-8], or with tri-*tert*-butylphosphine, it efficiently catalyses the cross-coupling of aryl- and heteroaryl-halides with aryl Grignard reagents [Böhm et al. *Angew Chem. Int Ed* 39 1602 2000].

**Nickel(II) bis(triphenylphosphine) dichloride** [**bis(triphenylphosphine)nickel(II) dichloride**] [14264-16-5] **M 654.2, m 247-250°(dec), 250°(dec)**. It is best prepared by adding  $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$  (2.38g, 10mmol) in  $\text{H}_2\text{O}$  (2ml) to a solution of  $\text{Ph}_3\text{P}$  (5.25g, 20mmol) in glacial AcOH (25ml) when an olive-green precipitate separates, but changes to dark blue crystals after standing in the mother liquors for 24 hours. These are filtered off washed with glacial AcOH and dried in a vacuum desiccator ( $\text{H}_2\text{SO}_4$  and KOH) until AcOH is removed to give 84% of dark blue crystals. [Venanzi *J Chem Soc* 719 1958, Cotton et al. *J Am Chem Soc* 83 344 1961, Kocienski et al. *J Org Chem* 54 1215 1989, *Beilstein* 16 IV 953]. With butyl chloride at 150-180° in a sealed tube followed by crystallisation from BuOH, it provides blue crystals with m 176° of  $[\text{Ni}(\text{BuCl})(\text{Ph}_3\text{P})_2\text{Cl}_2]$ . [Yamamoto *Bull Chem Soc Jpn* 27 501 1954, *Beilstein* 16 IV 953.]

In small amounts it catalyses the formation of terminal alkenes, e.g. from alkyl bromides or iodides in the presence of BuLi/DBU in THF at 25° in 50-80% yields [Jerapoulos & Smith *JCS Chem Commun* 1621 1986]. It also catalyses the replacement of OH groups in allyl alcohols with the alkyl group of alkylmagnesium bromide thus forming a C-C bond, and also causes dehydrohalogenation of alkylhalides to form the corresponding terminal olefin [Chuit et al. *JCS Chem Commun* 1604 1986].

**Nickel(0) bis(1,5-cyclooctadiene)** [**bis-(1,5-cyclooctadiene)nickel(0), Ni(COD)<sub>2</sub>**] [1295-35-8] **M 275.1, m 60°(dec), 142°(dec)**. It is available in sealed ampoules under  $\text{N}_2$ . All procedures should be carried out in a dry box

and in an atmosphere of N<sub>2</sub> or argon in subdued light because the complex is light and oxygen sensitive, and flammable. The solid is washed with dry Et<sub>2</sub>O (under argon) and separates from toluene as yellow crystals. Filter this under argon gas pressure, place the crystals in a container and dry them under a vacuum of 0.01mm to remove adhered toluene, flush with argon and seal them under argon or N<sub>2</sub> in glass ampoules. It catalyses cycloaddition reactions of 1,3-dienes [Semmelhack *Org Reactions* **19** 115 and 178 1972, Wilke et al. *Justus Liebigs Ann Chem* **699** 1 1966, Wender & Jenkins *J Am Chem Soc* **111** 6432 1989, Fieser & Fieser's *Reagents for Org Synth* **4** 33, 16 29, 17 32]. It also catalyses the addition of allyl phenyl sulfide to alkynes leading to 1,4-dienes. The reaction with acetylenes affords high yields, and in the presence of chiral phosphine ligands, with high stereoselectivity. The reaction tolerates a variety of functional groups [Hua et al. *Org Lett* **9** 263 2007].

**SUSPECTED CARCINOGEN.**

**Oxime Palladacycle dimers (Nájera Catalysts, OPs)** are Pd complexes of aromatic oximes, e.g. where palladium(II) is complexed with the oxime nitrogen atom and co-metalated with the *o*-position of the aromatic ring. They catalyse a variety of coupling reactions such as Suzuki-Miyaura, Heck, Stille, Sonogashira (all cross coupling) reactions and the Ullmann (homocoupling) reactions. *Note that the concentration of Pd in the products from Suzuki-Miyaura coupling reactions can be reduced from ~100ppm to 8000ppm by treating the reaction mixture with toluene and 20% aqueous NaHSO<sub>3</sub> at ~60° for ~1 hour* [Bullock, Mitchell and Toczko *Organic Process Research & Development* **12** 896 2008].

**Oximes** are prepared by a common procedure: A solution of the required aryl-ketone (22mmol) in MeOH (10ml) is added to a solution of hydroxylamine hydrochloride (3.06g, 44mmol), anhydrous NaOAc (3.6g, 44mmol) in H<sub>2</sub>O pre-heated at 60° for 1 hour. If some solid separates, add enough MeOH to obtain a clear solution, then stir at this temperature overnight. On cooling to ~25°, the oximes that separate in >90% yield are filtered off, washed with H<sub>2</sub>O and recrystallised to <sup>1</sup>H NMR purity (>98%). *Acetophenone oxime* has **m 59.1-59.7°** (from MeOH then cyclohexane, Pearson & Ball *J Org Chem* **14** 125 1949), *4,4'-dichlorobenzophenone oxime* has **m 135.2-136.9°** (from MeOH, Sieger & Klein *J Org Chem* **22** 953 1957), *4,4'-dihydroxybenzophenone oxime* has **m 266-267°dec.** (from EtOH, Zigeuner & Ziegler *Monatsh für Chemie* **89** 359 1949), *4-hydroxyacetophenone oxime* has **m 144-145°** (from aqueous MeOH), *4-methoxyacetophenone oxime* has **87°** (from petroleum ether, v. Auwers et al. *Chem Ber* **58** 41 1925), *4-methylacetophenone oxime* has **m 87-88°** (from MeOH then cyclohexane, Pearson & Ball *J Org Chem* **14** 125 1949), and *9-fluorenone oxime* has **m 195°, 198°** (yellow crystals from \*C<sub>6</sub>H<sub>6</sub> or xylene, Anet et al. *Can J Chem* **35** 180 1957, Wislicenus & Waldmuller *Chem Ber* **41** 3335 1908).

**Oxime palladacycles** are generally dimers that can be prepared by a general procedure. A solution of the oxime (10mmol) in methanol (20ml) containing NaOAc.3H<sub>2</sub>O (1.63g, 10mmol) is added to a 0.5M solution of Li<sub>2</sub>PdCl<sub>4</sub> (20ml, i.e. 2.62g, 10mmol, see [123334-21-4] in this Chapter Part 2) and stirred at ~25° for 2 to 3 days. Small quantities of precipitate may form and the colour of the reddish black solution turns to yellow. This is filtered and H<sub>2</sub>O (10ml) is added to precipitate the palladacycles as yellow solids (~90% yields) that are filtered off, washed with a little MeOH and H<sub>2</sub>O, and dried *in vacuo*. [Botella & Nájera *J Organomet Chem* **663** 46 2002, Onoue, Minami and Nakagawa *Bull Chem Soc Jpn* **43** 3480 1970; for preparation *via* ligand exchange see Ryabov et al. *Inorg Chem* **31** 3083 1992.] They can be used directly to catalyse reactions, their solubilities vary with the complex, but some can be crystallised from CH<sub>2</sub>Cl<sub>2</sub>-*n*-hexane, \*C<sub>6</sub>H<sub>6</sub>-*n*-hexane or CHCl<sub>3</sub>. They catalyse Heck couplings (in *N*-methylpyrrolidone with Et<sub>3</sub>N as base), Heck couplings under Jeffrey's conditions (DMF with Bu<sub>4</sub>NBr), Sonogashira reactions (with acetylenes in pyrrolidine and CuI), Stille coupling (in toluene), Suzuki coupling (in toluene with K<sub>2</sub>CO<sub>3</sub> as base), and Ullmann-type homocoupling (in DMF, in the necessary presence of hydroquinone), and typical reaction conditions are described [Alonso, Nájera and Pacheco *Org Lett* **2** 1823 2000, Iyer & Ramesh *Tetrahedron Lett* **41** 8981 2000].

**Di-chloro-bis[5-chloro-2-[(4-chlorophenyl)(hydroxyimino)methyl]phenyl-C]-di-palladium {Nájera Catalyst I, di- $\mu$ -chlorobis[5-chloro-2-[(4-chlorophenyl)(hydroxyimino- $\kappa$ N)methyl]phenyl- $\kappa$ C]-palladium (II) dimer}** [287410-78-0] has **M 814.0** and **m 208-210°**. [Alonso, Nájera and Pacheco *Org Lett* **2** 1823 2000].

**Di-chloro-bis[5-hydroxy-2-[1-(hydroxyimino)ethyl]phenyl-C]-di-palladium {Nájera Catalyst II, di- $\mu$ -chlorobis[5-chloro-2-[1-(hydroxyimino- $\kappa$ N)ethyl]phenyl- $\kappa$ C]-palladium(II) dimer}** [419581-64-9] has **M 584.1** and **m >250°, 251-255°**. It has IR (KBr) with  $\nu_{\max}$  at 3425 (OH), 1662 (C=N) cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, DMF-*d*<sub>7</sub>, TMS) has  $\delta_{\text{H}}$  at 9.96 (brs, 2H, 2 x OH), 9.71(brs, 2H, 2 x OH), 7.20 (brs, 2H, ArH), 7.07 (d, *J* = 8.5Hz, 2H, ArH), 6.50 (dd, *J* = 7.9, 2.4 Hz, 2H, ArH), 2.24 (s, 6H, 2 x CH<sub>3</sub>); and its <sup>13</sup>C NMR (75MHz, DMF-*d*<sub>7</sub>, TMS) has  $\delta_{\text{C}}$  at 167.3, 157.4, 152.7, 134.2, 127.2, 123.4, 111.3, 111.1. Botella & Nájera [*J Organomet*

*Chem* **663** 46 2002] obtained high turnover numbers and turnover frequencies in Suzuki-Miyaura cross-coupling reactions at room temperature conditions with this catalyst. Botella & Nájera [*J Org Chem* **70** 4360 2005] also studied this catalyst for Mirozoki-Heck couplings in aqueous *N,N*-dimethylacetamide in air using *N*-methyldicyclohexylamine (see [7560-83-0]) as base with or without Bu<sub>4</sub>NBr.

**Di-chloro-bis[2-[1-(hydroxyimino)ethyl]phenyl-C]-di-palladium {di-μ-chlorobis[2-[1-(hydroxyimino-κN)ethyl]phenyl-κC]-palladium(II) dimer}** [32679-19-9] has **M 568.1** and **m 210°**, **209-212°**. It has IR (KBr) with  $\nu_{\max}$  at 3426 (OH), 1640 (C=N) cm<sup>-1</sup>. [Alonso, Nájera and Pacheco *Org Lett* **2** 1823 2000, Onoue, Minami and Nakagawa *Bull Chem Soc Jpn* **43** 3480 1970.]

**Di-chloro-bis[5-hydroxy-2-[(4-hydroxyphenyl)(hydroxyimino)methyl]phenyl-C]-di-palladium, {di-μ-chlorobis[5-hydroxy-2-[(4-hydroxyphenyl)(hydroxyimino-κN)methyl]phenyl-κC]-palladium(II) dimer}** [419581-64-9] has **M 777.0** and **m >250°**. It has IR (KBr) with  $\nu_{\max}$  at 3405 (OH), 1612 (C=N) cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, DMF-*d*<sub>7</sub>, TMS) has  $\delta_{\text{H}}$  at 10.50-9.40 (brs, 6H, 6 x OH), 7.41 (d, *J* = 8.5Hz, 4H, ArH), 7.26 (brs, 2H, 2 x OH), 7.20 (brs, 2H, ArH), 7.02 (d, *J* = 8.5Hz, 4H, ArH), 6.69 (d, *J* = 8.5Hz, 4H, ArH), 6.48 (dd, *J* = 7.9, 2.7 Hz, 2H, ArH); and its <sup>13</sup>C NMR (75MHz, DMF-*d*<sub>7</sub>, TMS) has  $\delta_{\text{C}}$  at 167.8, 160.0, 157.5, 152.9, 134.3, 131.5, 129.4, 123.7, 121.0, 115.8, 111.5. [Botella & Nájera *J Organomet Chem* **663** 46 2002.]

**Di-chloro-bis[5-chloro-2-[(4-phenyl)(hydroxyimino)methyl]phenyl-C]-di-palladium, {di-μ-chlorobis[5-chloro-2-[(4-phenyl)(hydroxyimino-κN)methyl]phenyl-κC]-palladium(II) dimer}** [1145982-32-6; 30471-18-2 for stereoisomer] has **M 743.0** and **m 139-141°**. It has IR (KBr) with  $\nu_{\max}$  at 3373 (OH), 1645 (C=N), 1569, 1435, 1340, 1024 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, DMSO-*d*<sub>6</sub>, TMS) has  $\delta_{\text{H}}$  at 6.69 (m, 4H), 7.09 (m, 4H), 7.35-8.10 (m, 10H). [Alonso, Nájera and Pacheco *Org Lett* **2** 1823 2000, Onoue, Minami and Nakagawa *Bull Chem Soc Jpn* **43** 3480 1970.]

**Di-chloro-bis[5-methoxy-2-[(4-methoxyphenyl)(hydroxyimino)methyl]phenyl-C]-di-palladium {di-μ-chlorobis[5-methoxy-2-[(4-methoxyphenyl)(hydroxyimino-κN)methyl]phenyl-κC]-palladium(II) dimer}** [287410-79-1] has **M 803.0** and **m 135-137°**. It has IR (KBr) with  $\nu_{\max}$  at 3390 (OH), 1607 (C=N), 1579, 1559, 1253, 1233, 1177, 1026 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, DMSO-*d*<sub>6</sub>, TMS) has  $\delta_{\text{H}}$  at 3.71 (s, 6H), 3.83 (s, 6H), 6.71 (m, 4H), 7.08-7.25 (m with d at 7.10, *J* = 8.6Hz, 5H), 7.30-7.55 (m with d at 7.43, *J* = 8.6Hz, 5H). [Alonso, Nájera and Pacheco *Org Lett* **2** 1823 2000].

**Palladium(II) acetate [Pd(OAc)<sub>2</sub>]** [3375-31-3] **M 244.5**, **m 205°(dec.)**, **pK<sub>1</sub><sup>25</sup> 1.0**, **pK<sub>2</sub><sup>25</sup> 1.2 (for Pd<sub>2</sub><sup>2+</sup>)**. It recrystallises from CHCl<sub>3</sub> as purple crystals. It can be washed with AcOH and H<sub>2</sub>O and dried in air. Large crystals are obtained by dissolving it in \*C<sub>6</sub>H<sub>6</sub>, adding half its volume of AcOH and allowing it to evaporate slowly at room temperature. It forms green adducts with nitrogen donors, it dissolves in KI solution to form solid PdI<sub>2</sub> and a red solution of PdI<sub>4</sub><sup>2-</sup>, but is insoluble in aqueous saturated NaCl, and NaOAc. It dissolves in HCl to form PdCl<sub>4</sub><sup>2-</sup>. It is soluble in CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub>, Me<sub>2</sub>CO, MeCN, Et<sub>2</sub>O, but it is insoluble in H<sub>2</sub>O, and decomposes when warmed in alcohols in which it is also insoluble. [Morehouse et al. *Chem Ind (London)* 544 1964, Stephenson et al. *J Chem Soc* 3632 1965, Skapski & Smart *J Chem Soc (D)* 658 1970, Heck *Acc Chem Res* **12** 146 1979.] **Pd(OAc)<sub>2</sub>** should be an orange/red solid. Sometimes **commercial** samples, or old samples, give lower yields in catalytic reactions and it is found that good results are obtained after recrystallisation from \*C<sub>6</sub>H<sub>6</sub> [Anderson et al, *Org Synth* **84** 150 2007].

It is useful as a catalyst for coupling reactions [Kang et al. *J Am Chem Soc* **128** 6194 2006], and is a precursor for preparing heterogeneous and homogeneous catalysts [Jujjuri et al. *J Catal* **239** (2) 486 2006, Xu et al. *J Nano Res* **7** 449 2005].

**Palladium(II) acetylacetonate [Pd(acac)<sub>2</sub>]** [14024-61-4] **M 304.6**, **m 200-250°(dec)**. It can be recrystallised from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and sublimes *in vacuo*. It is soluble in heptane, \*C<sub>6</sub>H<sub>6</sub> (1.2% at 20°, 2.2 at 40°), toluene (0.56% at 20°, 1.4% at 40°) and acetylacetonate (1.2% at 20°, 0.05% at 40°). [West & Riley *J Inorg Nucl Chem* **5** 295 1957/8, Fernelius & Bryant *Inorg Synth* **V** 105 1957, *Beilstein* **1** IV 3676.] It is a soluble Pd source for the preparation of various soluble Pd catalysts by transferring the metal to a variety of phosphorus and other ligands for homogeneous catalysis [cf *Handbook of Organopalladium for Organic Synthesis* Negishi ed. Wiley, Hoboken NJ 2002, ISBN 0-471-31506-0.]

A complex consisting of Cu 1.10-phenanthroline [which mediates decarboxylation of arylcarboxylic acids with formation of aryl Cu species] and Pd(acac)<sub>2</sub> [for coupling] was made and used to catalyse the decarboxylative cross-coupling of the Cu species with aryl halides. This bimetallic system allows direct coupling of a variety of aryl, heteroaryl or vinyl carboxylic acids with aryl or heteroaryl bromides, chlorides or iodides at 160° in *N*-methylpyrrolidine in the presence of K<sub>2</sub>CO<sub>3</sub>. [Goossen et al. *J Am Chem Soc* **129** 4824 2007].

**Palladium EnCat.** Ley and co-workers [Ramarao et al. *J Chem Soc, Chem Commun* 1132 2002, Pears & Smith *Aldrichimica Acta* **38** 23 2005] have described a method of encapsulating Pd(OAc)<sub>2</sub> in the matrix of 20-250 μm microcapsules prepared from a dispersion of an aromatic polyfunctional isocyanate and Pd(OAc)<sub>2</sub> in CH<sub>2</sub>Cl<sub>2</sub> into H<sub>2</sub>O containing stabilisers and surfactants. At the point when the oily dispersion reaches the desired size, e.g. 20-250 μm, polymerisation is initiated (evolution of heat) whereby part of the isocyanate groups are hydrolyse to carbamate then to amino groups. The latter condense with the unhydrolysed isocyanate groups to form a crosslinked polyurea matrix entrapping the metal catalyst. After the necessary washing etc., polyurea microcapsules (MC average size ~150 μm) are formed which are hard, porous and highly crosslinked spheres. These beads are catalytically active, robust and recyclable and have been used in conventional and supercritical media (e.g. liquid CO<sub>2</sub>). They catalyse MC-[Pd] mediated Heck coupling (*p*-aryl and 4-heteroaryl nitro, methoxy and fluoro compounds with acrylic esters), carbonylation of 4-substituted arenes or heterocycles, Suzuki-type (ArB(OH)<sub>2</sub> + Ar to form crossed biaryls) and Stille couplings without requiring supplementary ligands [Ley et al. *J Chem Soc, Chem Commun* 1134 2002]. For **EnPd(OAc)<sub>2</sub>** [**Pd EnCat™**] and **Pd<sup>(0)</sup> EnCat™** see entries in Chapter 8.

Encapsulated Pd catalysts such as **Pd EnCat™** are available commercially. These particles are defined by their matrix content e.g. 30 or 40%, the latter having the smaller pore size. [See also Bremeyer et al. *Synlett* 1843 2002, Yu et al. *J C S Chem Commun* 678 2003, Vickerstaffe et al. *Org Biomol Chem* **1** 2419 2003.]

**Palladium(II) trifluoroacetate** [42196-31-6] **M 332.4, m ~210°(dec)**. Suspend it in trifluoroacetic acid and evaporate it on a steam bath a couple of times. The residue is then dried in vacuum (40-80°) to give a brown powder. It is *hygroscopic* and should be stored in a dry atmosphere, preferably aliquoted in sealed vials [Stephenson et al. *J Chem Soc* 3632 1965, Trost & Metzner *J Am Chem Soc* **102** 3572 1980.]

**Pd(CF<sub>3</sub>CO)<sub>2</sub>** catalyses the decarboxylation of electron-rich aromatic acids (e.g. with OMe groups) in DMSO/DMF at 70-90° (1 to 24 hours) in high yields and is not affected by steric hinderance [Dickstein et al. *Org Lett* **9** 2441 2007], and (in the presence of Cu(OAc)<sub>2</sub> with Cesium pivalate + 3-nitropyridine as additives at 110-140° in a microwave) it catalysed direct cross-coupling between unactivated arenes and *N*-acetylindoles with coupling mostly at C3 of indoles but with no homo-coupling [Stuart & Fangou *Science* **316** 1172 2007].

In the presence of AcOH, benzoquinone (as oxidant) and *o*-methoxyacetophenone or Ph<sub>3</sub>P as ligands, Pd(tfa)<sub>3</sub> catalyses selective allylic oxidation of olefins into their allyl acetates [McMurry & Kocovsky *Tetrahedron Lett* **25** 4187 1984].

**Phosferrox ligands and SK-Naud catalysts.** The ligands are (diphenylphosphinoferrrocenyl)oxazolines and are complexes where one of the cyclopentadienyl rings has two different substituents. These molecules have “**planar chirality**”, i.e. are asymmetric, and exist in two enantiomeric forms. The oxazolines are 4,5-dihydro-oxazoles which for this use have a 4-substituent, usually alkyl or aryl, thus introducing a chiral centre at C-4 of the heterocyclic ring. Phosferrox ligands are prepared by lithiation of chiral 4-alkyloxazolin-2-ylferrocene in which lithiation is directed predominantly to one of the *ipso* positions of the cyclopentadienyl ring, generating “**planar asymmetry**” in the ferrocene moiety. It is highly diastereoselective with *ortho*-lithiation yield of ~84-99% de, producing the *S*-stereochemistry at the ferrocene moiety as shown by X-ray analysis and CD spectra of the products. Reaction of the lithiated ferrocene with Ph<sub>2</sub>PCl furnishes the desired chiral phosferrox ligand. The stereochemistry of the reaction has been studied in detail [Sammakia et al. *J Org Chem* **60** 10 1995, Sammakia & Latham *J Org Chem* **60** 6002 1995, Richards & Mulvaney *Tetrahedron: Asym* **7** 1419 1996, Nishibayashi et al. *J Organomet Chem* **545-546** 381 1997]. Advantage is taken of the diastereoselectivity of the lithiation reaction in order to obtain the enantiomeric *R*-ferrocene. Thus, after lithiation, the lithium is displaced by a trimethylsilyl group (by reaction with Me<sub>3</sub>SiCl), the *S*-TMS derivative is lithiated again, but at the other *ipso* position of the same cyclopentadienyl ring, followed by reaction with Ph<sub>2</sub>PCl to form the 1-diphenylphosphino-3-TMS-2-(oxazolin-2-yl)ferrocene. Finally, removal of the TMS group, e.g. with *tetra-n-butylammonium fluoride* (TBAF, see [429-41-4; 3 H<sub>2</sub>O 87749-50-6]), provides the phosferrox where the stereochemistry at the ferrocene moiety is now *R*. Chiral phosferrox ligands react with RuCl<sub>2</sub>(Ph<sub>3</sub>P)<sub>3</sub> [15529-49-4] to form *Nauk catalysts* which are *phosferrox-RuCl<sub>2</sub>(Ph<sub>3</sub>P)* complexes that reduce aryl ketones in the presence of *i*-PrOH/*i*-PrOK to form the respective alcohols with high stereoselectivity [Sammakia & Stangeland *J Org Chem* **62** 6104 1997, Nishibayashi et al. *Organometallics* **18** 2291 1999].

**S-2-[(S)-2-(Diphenylphosphino)ferrocenyl]-4-(1-methylethyl)oxazoline (S,S-*i*-Pr-Phosferrox)** [163169-29-7] **M 481**, **m 157-158°**,  $[\alpha]_{\text{D}}^{24} +112^{\circ}$  (**c 0.1**, EtOH). This phosferrox is prepared in a Schlenk tube at  $-78^{\circ}$  under  $\text{N}_2$ , by adding dropwise *n*-BuLi (0.38ml, 0.7mmol) to a yellow-orange stirred solution of *S*-2-ferrocenyl-4-(1-methylethyl)oxazoline (0.158g, 0.53mmol) and TMEDA (0.10ml, 0.7mmol) in  $\text{Et}_2\text{O}$  (6ml) which had formed a yellow precipitate, and is stirred for 2 hours; the tube containing the orange non-homogeneous mixture is transferred to an ice-bath and stirred for 5 minutes further. To this now orange-red homogeneous solution is added  $\text{Ph}_2\text{PCl}$  (0.12ml, 0.7mmol, see [1079-66-9]), the mixture is allowed to warm to  $\sim 25^{\circ}$ , and after 15 minutes it is quenched with saturated aqueous  $\text{NaHCO}_3$  (10ml), the layers are separated, the aqueous layer is extracted with  $\text{Et}_2\text{O}$  (10ml), the combined  $\text{Et}_2\text{O}$  solutions are dried ( $\text{Mg SO}_4$ ), filtered, and evaporated to give an orange crystalline solid. This is purified by column chromatography (pre-adsorbed on silica, eluting with 10% EtOAc/petroleum ether) to give a yellow-orange crystalline solid (0.163g, 64%) that provided **S,S-*i*-Pr-Phosferrox** as an analytically pure single diastereomer upon recrystallisation from hexane. It has CD ( $\text{CHCl}_3$ )  $\lambda_{\text{max}}$  ( $\Delta\epsilon$ ) 456 (+2.20), 368 (+0.49), 342 (-1.00), 315 (+216) nm; the IR has  $\nu_{\text{max}}$  (nujol) at 1652 ( $\text{C}=\text{N}$ )  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (360Mz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 0.68 (3H, d,  $J = 7$  Hz,  $-\text{CH}_3$ ), 0.82 (3H, d,  $J = 7$  Hz,  $-\text{CH}_3$ ), 1.61-1.69 (1H, m,  $\text{CH}(\text{CH}_3)_2$ ), 3.61 (1H, brs, Fc), 3.67 (1H, t,  $J = 8$  Hz,  $-\text{OCHH}$ ), 3.83-3.90 (1H, m,  $-\text{NCH}$ -), 4.22 (5H, s,  $\text{C}_5\text{H}_5$ ), 4.22-4.30 (1H, m,  $-\text{OCHH}$ -), 4.37 (1H, brs, Fc), 4.99 (1H, brs, Fc), 7.18-7.24 (5H, m, Ph), 7.36-7.37 (3H, m, Ph), 7.46-7.51 (2H, m, Ph); the  $^{13}\text{C}$  NMR (90Mz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}\{^1\text{H}\}$  at 17.52 ( $-\text{CH}_3$ ), 18.61 ( $-\text{CH}_3$ ), 32.05 ( $\text{CH}(\text{CH}_3)_2$ ), 69.57 ( $-\text{OCH}_2$ -), 70.72 ( $\text{C}_5\text{H}_5$ ), 72.02, 72.14, 73.81, 73.85, 75.32 (d,  $J = 16$  Hz, Fc), 78.55 (d,  $J = 15$  Hz, Fc), 127.81 (Ph), 127.92 (Ph), 127.99 (Ph), 128.10 (Ph), 128.18 (Ph), 128.89 (Ph), 132.40 (d,  $J = 20$  Hz, Ph), 134.86 (d,  $J = 22$  Hz, Ph), 138.21 (d,  $J = 13$  Hz, Ph-*ipso*), 139.54 (d,  $J = 12$  Hz, Ph-*ipso*); the  $^{31}\text{P}$  NMR ( $\text{CDCl}_3$ ) has a single peak with  $\delta_{\text{P}}$  at -16.92, and the MS (EI) has  $m/z$  481 ( $\text{M}^+$ , 100%), 410 (68), 404 (44), 170 (38), 121 (76). [Richards & Mulvaney *Tetrahedron: Asym* 7 1419 1996].

**S-2-[(R)-2-(Diphenylphosphino)ferrocenyl]-4-(1-methylethyl)oxazoline [S-(R-*i*-Pr-Phosferrox)]** [163169-10-6] **M 481**, **m 132-132.5°**,  $[\alpha]_{\text{D}}^{20} -53^{\circ}$  (**c 0.15**, EtOH). This diastereomer is obtained by de-trimethylsilylation of *S*-2-[*R*-2-(diphenylphosphino)-5-(trimethylsilyl)ferrocenyl]-4-(1-methylethyl)oxazoline (0.14g 0.25mmol) with a yellow solution of 1M TBAF (see [429-41-4; 3  $\text{H}_2\text{O}$  87749-50-6], in Part 2) in THF (10ml) containing *ca* 5%  $\text{H}_2\text{O}$  by boiling for 4 hours, evaporating *in vacuo* to a small volume, shaking with  $\text{Et}_2\text{O}$  (10ml) and  $\text{H}_2\text{O}$  (10ml), separating, the aqueous layer is extracted with  $\text{Et}_2\text{O}$ , the ethereal layers are combined, dried ( $\text{Mg SO}_4$ ), filtered, evaporated *in vacuo*, and the residue is chromatographed in a silica column (eluted with 10% EtOAc/petroleum ether b 40-60°) to give **S-(R-*i*-Pr-Phosferrox)** as a yellow crystalline solid (0.090, 75%). Its CD ( $\text{CHCl}_3$ ) has  $\lambda_{\text{max}}$  ( $\Delta\epsilon$ ) 492 (+0.45), 434 (-0.43), 361 (-0.31), 344 (+0.35), 314 (-2.38) nm; the IR has  $\nu_{\text{max}}$  (nujol) at 1660 ( $\text{C}=\text{N}$ )  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (360Mz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 0.63 (3H, d,  $J = 7$  Hz,  $-\text{CH}_3$ ), 0.65 (3H, d,  $J = 7$  Hz,  $-\text{CH}_3$ ), 1.50-1.58 (1H, m,  $\text{CH}(\text{CH}_3)_2$ ), 3.62 (1H, brs, Fc), 3.89-4.08 (3H, m,  $-\text{OCH}_2\text{CH}$ -), 4.22 (5H, s,  $\text{C}_5\text{H}_5$ ), 4.36 (1H, brs, Fc), 4.94 (1H, brs, Fc), 7.19-7.23 (5H, m, Ph), 7.35-7.37 (3H, m, Ph), 7.48-7.53 (2H, m, Ph); the  $^{13}\text{C}$  NMR (90Mz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}\{^1\text{H}\}$  at 17.88 ( $\text{CH}_3$ ), 32.44 ( $-\text{CH}_3$ ), 69.47 ( $-\text{OCH}_2$ ), 70.44, 70.67 ( $\text{C}_5\text{H}_5$ ), 72.09, 72.41, 74.07, 74.95 (d,  $J = 16$  Hz, Fc), 78.12 (d,  $J = 14$  Hz, Fc), 127.71 (Ph), 127.83 (Ph), 127.90 (Ph), 128.07 (Ph), 128.15 (Ph), 128.55 (Ph), 128.86 (Ph), 132.38 (d,  $J = 19$  Hz, Ph), 134.98 (d,  $J = 22$  Hz, Ph), 138.31 (d,  $J = 13$  Hz, Ph-*ipso*), 139.37 (d,  $J = 12$  Hz, Ph-*ipso*), 164.62 ( $\text{C}=\text{N}$ ); the  $^{31}\text{P}$  NMR ( $\text{CDCl}_3$ ) has a single peak at  $\delta_{\text{P}}$  -18.03; and the MS (EI) has  $m/z$  481 ( $\text{M}^+$ , 32%), 410 (19), 170 (22), 121 (100). [Richards & Mulvaney *Tetrahedron: Asym* 7 1419 1996].

Among other phosferrox ligands that were prepared are *S,S*-4'*Me*-Phosferrox, *S,R*-4'*Me*-Phosferrox, *S,S*-4'*Et*-Phosferrox, *S,S*-4'*n*-Bu-Phosferrox, *S,S*-4'*iso*-Bu-Phosferrox, *S,S*-4'*t*-Bu-Phosferrox, *S,S*-4'*Ph*-Phosferrox, and *S,S*-4'*benzyl*-Phosferrox [Sammakia & Stangeland *J Org Chem* 62 6104 1997, Nishibayashi et al. *J Organomet Chem* 545-546 381 1997].

**S-2-[*S*<sub>p</sub>-2-(Diphenylphosphino)ferrocenyl]-4-*iso*-propyl-2-oxazoline triphenylphosphine ruthenium (II) dichloride complex [S,S-*i*-Pr-Phosferrox-Ru(II)  $\text{Ph}_3\text{PCl}_2$  complex, SK-Naud Catalyst-N003-2z]** [212133-11-4] **M 915.6**,  $[\alpha]_{\text{D}}^{20} -1100^{\circ}$  (**c 0.2**,  $\text{CHCl}_3$ ). This catalyst can be prepared *in situ* from the phosferrox and  $\text{RuCl}_2(\text{PhP}_3)_3$ , but it can also be isolated. Thus a mixture of  $\text{RuCl}_2(\text{PhP}_3)_3$  (480mg, 0.50mmol) and *S,S*-*i*-Pr-Phosferrox (0.50mmol) and toluene (15ml) are stirred under  $\text{N}_2$  at  $\sim 25^{\circ}$  for 20 hours when the original purple solution changes to a red suspension. Addition of *n*-hexane caused the crystalline catalyst to separate, and the red crystals are recrystallised to analytical purity from  $\text{CH}_2\text{Cl}_2$ /*n*-hexane ( $\sim 81\%$  yield). The  $^1\text{H}$  NMR (270Mz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 0.57 (3H, d,  $J = 7$  Hz,  $\text{CH}_3$ ), 0.97 (3H, d,  $J = 7$  Hz,  $\text{CH}_3$ ), 2.12 (1H, dd,  $J = 8$  and 8Hz), 3.21 (1H, m), 3.28 (1H, m), 3.80 (1H, dd,  $J = 3$  and 8Hz), 4.02 (5H, s), 4.59 (1H, m), 4.68 (1H, m), 4.84 (1H, m), 6.5-8.4 (25H, m); and the  $^{31}\text{P}$  NMR (107Mz,  $\text{CDCl}_3$ ) indicated a diastereomerically pure complex with  $\delta_{\text{P}}$  40.1 (d,  $J =$

45Hz) and 77.0 (d,  $J = 45\text{Hz}$ ). It causes a high conversion with high stereoselectivity (with 88- $<99\%$  ee) in the Ru-catalysed asymmetric transfer hydrogenation of ketones using *i*-PrOH as H source, as well as catalysing the asymmetric oxidation of *sec*-alcohols to ketones [Nishibayashi et al. *Organometallics* **18** 2291 1999].

**R-2-[R<sub>p</sub>-2-(Diphenylphosphino)ferrocenyl]-4-*iso*-propyl-2-oxazoline triphenylphosphine ruthenium(II) dichloride complex (SK-Naud Catalyst-N003-1z) [849921-25-1] M 915.6,  $[\alpha]_{\text{D}}^{20} +1100^{\circ}$  (c 0.5, CHCl<sub>3</sub>)** can be prepared and purified as in the preceding entry.

**S-2-[S<sub>p</sub>-2-(Diphenylphosphino)ferrocenyl]-4-phenyl-2-oxazoline triphenylphosphine ruthenium (II) dichloride complex [S,S-*Ph*-Phosferrox-Ru(II) Ph<sub>3</sub>P Cl<sub>2</sub> complex] [212210-73-6] M 949.6.** As above (SK-Naud catalyst-N003-2z) the catalyst can be prepared *in situ*, but can also be isolated in pure crystalline form, and its structure has been confirmed by X-ray crystallography. A mixture of **S,S-4'*Ph*-Phosferrox** [257mg, 0.50mol, red solid purified by flash chromatography with silica gel and eluting with hexanes the 10:1 hexanes/EtOAc, and recrystallisation from hexane, **m 184-185° dec.**,  $[\alpha]_{\text{D}}^{23} +25.6^{\circ}$  (c 0.67, CHCl<sub>3</sub>)] and RuCl<sub>2</sub>(PhP<sub>3</sub>)<sub>3</sub> (480mg, 0.50mmol) in toluene (15ml) are stirred under N<sub>2</sub> at  $-25^{\circ}$  for 20 hours when the original purple solution changes to a red suspension. Addition of *n*-hexane causes the crystalline catalyst to separate, and the red crystals are recrystallised to analytical purity from CH<sub>2</sub>Cl<sub>2</sub>/*n*-hexane to provide the catalyst (with 1 mol of CH<sub>2</sub>Cl<sub>2</sub>) in 80% yield. The <sup>1</sup>H NMR (270Mz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 2.68 (1H, dd  $J = 8$  and 9Hz), 3.29 (1H, dd,  $J = 1$  and 9Hz), 4.12 (5H, s), 4.23 (1H, dd,  $J = 1$  and 8Hz), 4.71 (1H, m), 4.80 (1H, m), 5.04 (1H, m), 6.6-8.2 (30H, m); and the <sup>31</sup>P NMR (107Mz, CDCl<sub>3</sub>) indicated almost diastereomerically pure complex with  $\delta_{\text{P}}$  40.8 (d,  $J = 45\text{Hz}$ ) and 75.7 (d,  $J = 45\text{Hz}$ ). It has similar catalytic activity as the SK-Naud catalysts [Nishibayashi et al. *Organometallics* **18** 2291 1999, Sammakia & Stangeland *J Org Chem* **62** 6104 1997].

**Platinum(0) bis(1,5-cyclooctadiene) [Pt(COD)<sub>2</sub>, Pt bis(1,2,5,6- $\eta$ -1,5-octadiene)] [12130-66-4] M 411.6, **m** $\sim$ **200°**. Pt(COD)<sub>2</sub> is prepared in a dry, O<sub>2</sub>-free, N<sub>2</sub> atmosphere by adding dropwise an ethereal solution of Li<sub>2</sub>(C<sub>8</sub>H<sub>8</sub>) [*ca* 40ml of a 0.24mol dm<sup>-3</sup> of dilithium cyclooctatetraene [40698-91-7] in Et<sub>2</sub>O (Katz *J Am Chem Soc* **82** 3784 1960, see also below)] to a finely powdered suspension of Pt(COD)Cl<sub>2</sub> (3.7g, 10mmol, see above entry) in 1,5-cyclooctadiene at  $-30^{\circ}$ , stirring for 30 minutes as the temperature rises to  $-10^{\circ}$ ; and then the solvent is evaporated *in vacuo* to dryness. Extraction of the residue with toluene (6 x 50ml) gives a brown solution which is filtered through an Al<sub>2</sub>O<sub>3</sub> column (8 x 2.5cm, Brockman Activity III), the filtrate is evaporated to  $\sim$ 15ml and the supernatant is decanted from the off-white product which is washed with several small volumes of cold toluene until the washings are almost colourless. This Pt catalyst (1.6-2.4g, 40-60%) is useful for most purposes. Analytically pure (C and H) white crystals can be obtained by dissolving it in petroleum ether (b 40°-60°) ( $\sim$ 80ml for each mmol of catalyst), filtering through an Al<sub>2</sub>O<sub>3</sub> column (6 x 2.5cm) and cooling to  $-78^{\circ}$ . Its IR has  $\nu_{\text{max}}$  at 3020m, 3000sh, 2930m, 2910sh, 2965m, 2825m, 1475m, 1465m, 1433m, 1327s, 1306w, 1238m, 1212m, 1178w, 1158w, 1090vw, 1075w, 1009w, 998m, 970sh, 962w, 890vw, 865sh, 855m, 829sh, 823m, 810m, 788m, 769vw, 580vw, 510vw, and 465vw cm<sup>-1</sup>; the <sup>1</sup>H NMR (100MHz, \*C<sub>6</sub>D<sub>6</sub>) has  $\tau$  at 5.80 (m, 8H, CH=CH,  $J_{\text{Pt-H}} = 55\text{Hz}$ ) and 7.81 (m, 16H, CH<sub>2</sub>); and the <sup>13</sup>C NMR (25.15MHz, \*C<sub>6</sub>D<sub>6</sub>-<sup>1</sup>H decoupled) has <sup>13</sup>C shifts with  $\delta$  measured relative to SiMe<sub>4</sub> (positive values to high frequency) 73.3 (C=C,  $J_{\text{Pt-C}} = 143\text{Hz}$ ) and 33.2 (CH<sub>2</sub>,  $J_{\text{Pt-C}} = 15\text{Hz}$ ). The white crystalline complex is stable in air and can be handled without difficulty. [Spencer *Inorg Synth* **19** 214 1979, Crascall & Spencer *Inorg Synth* **28** 127 1990.] Note that attempts to prepare the corresponding Pd complex [Pd(COD)<sub>2</sub>] in a similar manner failed; however, it was obtained by the reaction of [Pd(COD)Cl<sub>2</sub>] with Li<sub>2</sub>[C<sub>8</sub>H<sub>8</sub>] in the presence of excess of propene as a white crystalline solid which was stable below  $-20^{\circ}$ , but decomposed rapidly to Pd and 1,5-cyclooctadiene at ambient temperatures. [Green et al. *J Chem Soc, Dalton Trans* 271 1977, Crascall & Spencer *Inorg Synth* **28** 128 1990.] Pt(COD)<sub>2</sub> catalyses the *cis*-diborylation of olefins in  $>80\%$  yields under ambient conditions with e.g. CatB-BCat (see below). [Iverson & Smith *Organometallics* **16** 2757 1997].**

**Platinum(0) bis(dibenzylideneacetone) [Pt(dba)<sub>2</sub>, bis(C<sup>3</sup>,O<sup>3</sup>- $\eta$ -1,5-diphenyl-1,4-pentadien-3-one)-platinum (0)] [33677-56-4] M 683.7, **m** $>$ **170° (to Pt and dba)**. Pt(dba)<sub>2</sub> is a deep purple air stable complex which is prepared in 41% yield from K<sub>2</sub>PtCl<sub>4</sub> with 3 mol equivalents of dba in refluxing EtOH under N<sub>2</sub> in the presence of NaOAc. It crystallises from Me<sub>2</sub>CO and is soluble in MeOH and CHCl<sub>3</sub>. It is useful for preparing complexes with (PPh<sub>3</sub>)<sub>2</sub>, (AsPh<sub>3</sub>)<sub>2</sub>, (PEt<sub>3</sub>)<sub>2</sub>, (AsEt<sub>3</sub>)<sub>2</sub>, as well as with other ligands. The *tris* complex Pt(dba)<sub>3</sub> is formed as a yellow microcrystalline solid in 40% yield by carrying out the above reaction in refluxing aqueous MeOH *in the presence of air or oxygen*. It is stable in the solid form, and is soluble in organic solvents but gives green solutions that rapidly turns purple, from which Pt(dba)<sub>2</sub> and dba (1:1) can be isolated. Physical properties showed**

that complete dissociation of  $\text{Pt}(\text{dba})_3$  had occurred. Both complexes decompose  $>170^\circ$  on heating to Pt and dba, on treatment with CO, or when solutions are set aside for long periods. The IR ( $\text{CHCl}_3$ ) of  $\text{Pt}(\text{dba})_2$  has  $\nu_{\text{max}}$  at 1656w(CO), 1613vs, br(C=C), 1579m(C=C aromatic), 1544m, br(CO)  $\text{cm}^{-1}$ ; and complex  $\text{Pt}(\text{dba})_3$  has  $\nu_{\text{max}}$  at 1652w(CO), 1624vs(C=C), 1593s, br(C=C aromatic), 1579m(C=C aromatic), 1527m br(CO)  $\text{cm}^{-1}$ . The UV-VIS ( $\text{CHCl}_3$ ) of  $\text{Pt}(\text{dba})_2$  has  $\lambda_{\text{max}}$  nm ( $\log \epsilon$ ) at 235 (4.48), 336 (4.54), 378sh(4.27) 538 (3.92, metal  $\rightarrow$  ligand:  $d \rightarrow \pi^*$ ); of  $\text{Pt}(\text{dba})_3$  has  $\lambda_{\text{max}}$  nm ( $\log \epsilon$ ) at 235 (4.27), 336 (4.48), 568 (3.51, metal  $\rightarrow$  ligand:  $d \rightarrow \pi^*$ ); and of dba has  $\lambda_{\text{max}}$  nm ( $\log \epsilon$ ) at 233 (3.90), 325 (4.38). The  $^1\text{H}$  NMR (100MHz,  $\text{CDCl}_3$ ) of  $\text{Pt}(\text{dba})_2$  and dba are surprisingly similar [see however: Müller & Gröser *Angew Chem. Int Edn* **6** 364 1967]. It is a useful catalyst for the oligomerisation of acetylenes [Moseley & Maitlis *J C S. Chem Commun* 982 1971, Cherwinski et al. *JCS, Dalton Trans* 1405 1974], and is used in the  $\text{Pt}(\text{dba})_2$ -catalysed selective *cis*-addition of bis(pinacolato)borane to terminal alkenes and cyclic alkenes with internal strain to form bis(boryl)alkanes in 76-86% yields at  $50^\circ$  [Ishiyama et al. *J C S, Chem Commun* 689 1997] among other uses in catalysis.

**Rhodium(II) acetate dimer ( $2\text{H}_2\text{O}$ ) [dirhodium tetraacetate, tetrakis(acetato)dirhodium(II)] [15956-28-2] M 478.0.** Dissolve 5g of the dimeric salt in boiling MeOH (*ca* 600ml) and filter. Concentrate the filtrate to 400ml and chill overnight at *ca*  $0^\circ$  to obtain dark green crystals of the MeOH adduct. Concentration of the mother liquors gives a further crop of  $[\text{Rh}(\text{OAc})_2]_2 \cdot 2\text{MeOH}$ . The adduct is then heated at  $45^\circ$  in a vacuum for 2 hours (all MeOH is lost) to leave the emerald green crystals of the acetate. [Legzdins et al. *J Chem Soc (A)* 3322 1970, Rempel et al. *Inorg Synth XIII* 90 1972.] Alternatively, dissolve the acetate in glacial AcOH and reflux for a few hours to give an emerald green solution. Evaporate most of the AcOH on a steam bath, then heat the residue at  $120^\circ/1$  hour. Extract the residue with boiling  $\text{Me}_2\text{CO}$ . Filter, concentrate to half its volume and keep at  $0^\circ/18$  hours. Collect the crystals, wash them with ice cold  $\text{Me}_2\text{CO}$  and dry them at  $110^\circ$ . It is moderately soluble  $\text{H}_2\text{O}$ , MeOH,  $\text{Me}_2\text{CO}$  (see above), and in many organic solvents to give green solutions. It forms adducts with MeOH,  $\text{Me}_3\text{N}$  and  $\text{Me}_2\text{S}$ , and gives solutions with different colours varying from green to orange and red depending on solvent and dilution. The IR(Nujol) has  $\nu_{\text{max}}$  at 1580s, 1425s, and 1350m  $\text{cm}^{-1}$ , and the IR (hexachlorobutadiene) has  $\nu_{\text{max}}$  at 1445s, 1415s and 1350m  $\text{cm}^{-1}$  among other bands. [UV: Johnson et al. *Inorg Chem* **2** 960 1963, Beilstein **1** H 124.]

It is a homogeneous catalyst [Black *Aldrichimica Acta* **15** 13 1982], is used in an efficient synthesis for  $\beta$ -hydroxy- $\alpha$ -acrylates involving the decomposition of diazoester intermediates with concomitant 1,2-arylmigration [Xiao et al. *Tetrahedron Lett* **48** 1147 2007], and is an effective catalyst for the formation of allylsulfonium ylides (e.g. from 3,3-dimethylallyl methyl sulfide and trimethylsilyldiazomethane) which undergo [2.3] sigmatropic rearrangement (e.g. to 2,2-dimethyl-1-methylthio-1-trimethylsilylbut-3,4-ene) [Carter & Van Vranken *Tetrahedron Lett* **40** 1617 1999]. The reaction of alkyl diazoacetates with carbodiimides (to form 2-imino-4-oxazolines) [Drapier et al. *Tetrahedron Lett* 559 1979], with acetylenes (to form cyclopropenes) [Petiniot et al. *Tetrahedron Lett* 1239 1978], and with ROH, HOH and  $\text{R}'\text{COOH}$  (replacing H to form the respective  $\text{O}-\text{CH}_2\text{CO}_2\text{Alkyl}$ ) [Paulissen et al. *Tetrahedron Lett* 2233 1973]; are all catalysed by  $[\text{Rh}(\text{OAc})_2]_2$  at about room temperature.

**Rhodium(III) acetylacetonate [2,4-pentanedione rhodium(III),  $\text{Rh}(\text{acac})_3$ ] [14284-92-5] M 400.2, m 263-264 $^\circ$ .** It is prepared from  $\text{Rh}(\text{NO}_3)_3$  solution (0.1g in Rh) in 0.2N  $\text{HNO}_3$  (10ml) which is neutralised with aqueous  $\text{NaHCO}_3$  (10%) to pH 4, whereby the light yellow hydroxide (or basic nitrate) begins to separate. Acetylacetonone (5ml) is added, the mixture is refluxed, and after a few minutes orange-yellow crystals begin to separate while the pH of the solution decreases. After 30 minutes reflux the pH of the solution is re-adjusted and reflux is continued for a further 15 minutes. The orange-yellow complex is collected and forms monoclinic plates (0.3g, 75%, m  $260^\circ$ ) upon recrystallisation from aqueous MeOH, and sublimes at  $240^\circ/1.0\text{mm}$ . It decomposes above  $280^\circ$  depositing a rhodium mirror. It is insoluble in  $\text{H}_2\text{O}$ , slightly soluble in EtOH and petroleum ether, but freely soluble in  $^*\text{C}_6\text{H}_6$  and  $\text{CHCl}_3$ . It is stable in boiling dilute acids and 10% aqueous NaOH. Molecular weight determination ( $\sim 490$ , by Rast in camphor) indicates that it is monomeric. [Dwyer & Sargeson *J Am Chem Soc* **75** 984 1953, Beilstein **1** IV 3677.]  $\text{Rh}(\text{acac})_3$  is a very effective catalyst ( $\sim 90\%$  yields) for the hydrogenation of monocarboxylic or  $\omega$ -dicarboxylic acids to the respective alcohols in DME at 100 atmospheres (16 hours at  $160^\circ$ ) when combined with  $\text{Re}(\text{CO})_{10}$  or  $\text{Mo}(\text{CO})_6$  [He et al. *Tetrahedron Lett* **36** 1059 1995].

**Rhodium(I) bis(1,5-cyclooctadiene) tetrafluoroborate complex [bis(1,5-cyclooctadiene)rhodium(I) tetra-**

**fluoroborate, Rh(COD)<sub>2</sub><sup>+</sup>.BF<sub>4</sub><sup>-</sup> [35138-22-8] M 406.1, m ~190° (dec), the xHydrate [207124-65-0] has m 165° (dec).** This complex is prepared from the dimer [Rh(COD)Cl]<sub>2</sub> (1.47g, 2.98mmol, cf [12092-47-6]) in CH<sub>2</sub>Cl<sub>2</sub> (20ml), to which is added COD (1.1ml, 8.97mmol) followed by AgBF<sub>4</sub> (1.33g, 6.83mmol) in Me<sub>2</sub>CO (10ml) which results in a deep red solution containing a white precipitate. The mixture is stirred for 20 minutes, filtered through Celite, THF (20ml) is then added and the volume is reduced on a Rotovap at 25° down to 10ml. The deep red crystals are filtered off, washed with THF (2 x 5ml), Et<sub>2</sub>O and dried in air (2.35g, 97%). [Schenck et al. *Inorg Chem* **24** 2334 1985.] A general purification procedure involves dissolving the complex in the minimum volume of CH<sub>2</sub>Cl<sub>2</sub>, adding an equal volume of EtOH and completing the crystallisation by dropwise addition of Et<sub>2</sub>O, filtration, and drying the solid *in vacuo*.

It is used for preparing cationic COD-Rh complexes with phosphine ligands for enantioselective [2+2+2] cycloaddition of unsymmetrical diynes with styrene and norbornene derivatives to yield bi- and tetra-cyclic products with good (~50%) to very good (>90%) enantiomeric enrichment [Shibata et al. *Tetrahedron* **63** 12853 2007], for hydrogenation [Nagel et al. *Chem Ber* **119** 3326 1986, Ojima et al. *Tetrahedron* **45** 6901 1989, Sawamura et al. *J Am Chem Soc* **117** 9602 1995], and for hydrosilylation [Takeuchi et al. *J Org Chem* **60** 3045 1995].

**[Rhodium(S,S-Chiraphos)(COD)] ClO<sub>4</sub>.THF {(1,2,5,6-η)-1,5-cyclooctadiene}[1,2-dimethyl-1,2-ethanediyl-(di-phenylphosphine)-P,P']rhodium(1<sup>+</sup>) perchlorate. THF, η<sup>4</sup>-1,5-cyclooctadiene[(2S,3S)-2,3-bis(diphenylphosphino)butane]rhodium(I) perchlorate.THF} [61886-03-1, 61886-02-0 THF-free] M 809.1, 737.0 (THF-free).** The S,S-catalyst is obtained by adding 70% perchloric acid (0.080g, 1 equivalent) in pure THF (1ml) under N<sub>2</sub> to a mixture of S,S-chiraphos (0.244g, see [64896-28-2]) and \*[Rh(COD)(acac)] (0.180) in THF (4ml); and the red mixture is allowed to stand at 25° for 14 hours. The bright orange block crystals are collected, washed with cold THF and dried in air to give analytically pure (C, H, P and Cl analysis) catalyst *perchlorate THF* (0.40g). The presence of solvent is confirmed by NMR. A single crystal X-ray (absolute) structural determination of this THF-pre-catalyst shows that it is as predicted, i.e. the chiral centres are S and the methyl groups are equatorial to give a δ-chelate ring [Ball & Payne *Inorg Chem* **16** 1187 1977]. The configuration in solution is assumed to be the same, being consistent with the stereochemistry of the catalytically produced products.

The above two complexes are efficient homogeneous catalysts for the hydrogenation of α-N-acylamioacrylic acids at room temperature and pressure in THF, dioxane or \*C<sub>6</sub>H<sub>6</sub>, EtOH or aqueous EtOH to provide the amino acid derivatives in very high optical purity [e.g. of alanine (91%), DOPA (83%) and tyrosine (92%)], and in almost quantitative chemical yields with turnover numbers ranging from 3 x 10<sup>-2</sup> to 6 x 10<sup>-4</sup> sec<sup>-1</sup>, from which essentially completely optically pure “non-natural” R-α-amino acids can be obtained by recrystallisation. [Fryzuk & Bosnich *J Am Chem Soc* **99** 6262 1977]. An ingenious application of this S,S-chiraphos catalyst using hydrogen, deuterium and tritium has been adopted for the asymmetric synthesis of chiral lactic acid in which the methyl group is chiral by virtue of its having a hydrogen, a deuterium and a tritium atom on the methyl carbon atom. [Fryzuk & Bosnich *J Am Chem Soc* **101** 3043 1979]. In the hydrogenation of selected Z-olefins with this pre-catalyst, addition of Et<sub>3</sub>N appears to give higher optical yields of products [Onuma et al. *Bull Chem Soc, Jpn* **53** 2012 1980].

\* **[Rh(COD)(acac)] [12245-39-5] M 310.2, m 138-140°**, is prepared from [Rh(COD)Cl]<sub>2</sub> (4.4g, 9 mmol, see [12092-47-6]) and acetylacetone (acacH, 0.9ml, 9 mmol) in CH<sub>2</sub>Cl<sub>2</sub> (50ml) chilled to -80° by treating dropwise with a solution of Na<sub>2</sub>CO<sub>3</sub> (5.7g) in H<sub>2</sub>O (100ml) and shaking; then warming to 0° while shaking and more CH<sub>2</sub>Cl<sub>2</sub> (50ml) is added. The layers are separated, and the organic layer is evaporated *in vacuo* until the complex crystallises out. [Fryzuk & Bosnich *J Am Chem Soc* **99** 6262 1977].

**Rhodium(S,S-Chiraphos)(NBD)] ClO<sub>4</sub> {(bicyclo[2.2.1]heptadiene)[(2S,3S)-2,3-bis(diphenylphosphino)butane]-rhodium (I) perchlorate, [(1,2,5,6-η)bicyclo[2.2.1]hepta-2,5-diene][S,S-1,2-dimethyl-1,2-ethanediyl)-bis(diphenylphosphine)-P,P']rhodium(1<sup>+</sup>) perchlorate} [65012-74-0; 65012-73-9 for cation: Rh(S,S-chiraphos)(MeOH)<sub>2</sub><sup>+</sup> has 71264-71-6] M 720.97.** This catalyst can be prepared by two methods. In the first, 7% HClO<sub>4</sub> (0.173g, 1 equivalent) in THF (6ml) is added to a solution of the complex \*[Rh(NBD)(acac)] (0.356g) and S,S-chiraphos (0.516g, see [64896-28-2]) in dry, freshly distilled, THF (7.5ml) under N<sub>2</sub>, and the deep red coloured solution is allowed to stand at 25° for 24 hours. The orange-red crystals that separate are collected, washed with cold THF, dried at 40° *in vacuo* to give the catalyst *perchlorate* (0.6g), which should be

stored at 0° under N<sub>2</sub> in a sealed container. In the *second*, *S,S*-chiraphos (0.308g) and the complex **\*\*[Rh(NBD)<sub>2</sub>]ClO<sub>4</sub>** (0.290g) are dissolved in CH<sub>2</sub>Cl<sub>2</sub> (5ml) and THF (5ml) under N<sub>2</sub>, followed by addition of hexane (6ml), allowing to stand at 25° for 1 hour, then at 5° for 2 hours. The orange-red needles of the catalyst perchlorate (0.43g) are collected, dried and stored as in the first method. Note that it loses catalytic activity if stored in air. [Fryzuk & Bosnich *J Am Chem Soc* **99** 6262 1977.] Its <sup>31</sup>P NMR (CDCl<sub>3</sub>, with external H<sub>3</sub>PO<sub>4</sub> as reference) has δ at +56.9 (d, *J*<sub>Rh-P</sub> = 153Hz), i.e. downfield from H<sub>3</sub>PO<sub>4</sub> [Slack et al. *Inorg Chem* **18** 3125 1979]. The tetrafluoroborate **[79790-89-9] M 708.3** forms orange-red crystals from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O.

**\*[Rh(NBD)(acac)]** [32354-50-0] **M 294.2, m 176-177°**, is prepared from the dimer [Rh(NBD)Cl]<sub>2</sub> [12257-42-0] and Ti(acetylacetonate) ([25955-51-6]) as described for [(hfacac)(COE)<sub>2</sub>Ir] [58616-58-3] [see this chapter and Diversi et al. *J Organomet Chem* **125** 253 1977.]

**\*\* [Rh(NBD)<sub>2</sub>]<sup>+</sup>ClO<sub>4</sub><sup>-</sup>** [60576-59-1] **M 386.6**, is also prepared from the dimer [Rh(NBD)Cl]<sub>2</sub> (0.35g, [12257-42-0]) and NBD (0.14g) in CH<sub>2</sub>Cl<sub>2</sub> (15ml) under N<sub>2</sub>, adding AgClO<sub>4</sub> (0.315g), stirring for 1 hour, filtering and adding THF (15ml). When the CH<sub>2</sub>Cl<sub>2</sub> in the mixture is evaporated off under a vacuum, [Rh(NBD)<sub>2</sub>]ClO<sub>4</sub> separates as orange needles which are collected, washed with cold THF and dried *in vacuo* to give now orange-brown crystals (0.5g). *Alternatively*, add Et<sub>2</sub>O instead of THF (15ml), filter, wash with Et<sub>2</sub>O, and dry *in vacuo* (yield 90%). Its <sup>1</sup>H NMR in (CD<sub>3</sub>)<sub>2</sub>CO has δ at 5.23 (br, olefin H), 4.09 (bridgehead CH) and 1.50 (CH<sub>2</sub>). [Fryzuk & Bosnich *J Am Chem Soc* **99** 6262 1977; cf Cramer *J Am Chem Soc* **86** 217 1964, Uson et al. *Trans Metallorg Chem* **4** 55 1979.] **[Rh(NBD)<sub>2</sub>]<sup>+</sup>BF<sub>4</sub><sup>-</sup>** has [36620-11-8] **M 373.99, m 157-159°**, and is prepared as for the perchlorate, but replacing AgClO<sub>4</sub> by an equivalent of AgBF<sub>4</sub>, and recrystallising from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O. The cation **[Rh(NBD)<sub>2</sub>]<sup>+</sup>** has [35015-46-4] **M 287.2**. [Green et al. *J Chem Soc (A)* 2334 1971, Osborn et al. *J Am Chem Soc* **93** 3089 1971, Green et al. *JCS, Dalton Trans* 832 1972.]

**[Rhodium(*R,R*-Chiraphos)(NBD)]ClO<sub>4</sub> {(bicyclo[2.2.1]heptadiene)[(2*R*,3*R*)-2,3-bis(diphenylphosphino)butane]-rhodium (I) perchlorate}** [74892-62-9] **M 720.97**. This pre-catalyst is prepared and purified as its enantiomer above in 77% yield in orange-red crystals after recrystallising from CH<sub>2</sub>Cl<sub>2</sub>/hexane. Its catalytic properties are very similar to those of its enantiomer above except that optically pure “natural” *S*-α-amino acids are produced. [Köiiner & Gerber *Chem Ber* **113** 2323 1980.]

The mechanism of the above asymmetric hydrogenations has been studied in some detail [Brown & Chaloner *J Am Chem Soc* **102** 3040 1980, Brown et al. *J Organomet Chem* **216** 263 1981].

**[Rhodium (*R*-Prophos)(NBD)] ClO<sub>4</sub>. 0.5CH<sub>2</sub>Cl<sub>2</sub> {(bicyclo[2.2.1]heptadiene)[*R*(+)-1,2-bis(diphenylphosphino)butane]-rhodium(I) perchlorate. 0.5CH<sub>2</sub>Cl<sub>2</sub>, [(2,3,5,6-η)bicyclo[2.2.1]hepta-2,5-diene][(1-methyl-1,2-ethanediyl)bis(diphenylphosphine)-*P,P'*]rhodium(I<sup>+</sup>) perchlorate. 0.5CH<sub>2</sub>Cl<sub>2</sub>}** [67881-59-8; 67884-58-7 for the cation; *Rh*(*R*-prophos)(MeOH)<sub>2</sub><sup>+</sup> has 71264-72-7] **M 706.9, decomposes on heating**. The precatalyst is made from freshly recrystallised [Rh(NBD)<sub>2</sub>]ClO<sub>4</sub> (0.388g, 1mmol, [12257-42-0]) and *R*-prophos (0.437g, 1.106mmol) in a mixture of CH<sub>2</sub>Cl<sub>2</sub> (4ml) and pure THF (4ml) under N<sub>2</sub>; and to this orange red solution is added hexane (4ml) dropwise, and the mixture is then allowed to stand at 25° for 5 hours then at 5° for 12 hours.

The orange-red solid is filtered off quickly, washed with ice-cold THF then hexane, and dried under a stream of dry N<sub>2</sub> to give analytically pure [Rh(*R*-Prophos)(NBD)]ClO<sub>4</sub> (with 0.5CH<sub>2</sub>Cl<sub>2</sub> by NMR). It remains catalytically active indefinitely if kept at 0° under N<sub>2</sub>. Its <sup>31</sup>P NMR (CDCl<sub>3</sub>, with external H<sub>3</sub>PO<sub>4</sub> as reference) has δ at +60.5 (d, *J*<sub>Rh-P</sub> = 172Hz), +41.8 (q, *J*<sub>Rh-P</sub> = 139Hz, *J*<sub>P-P</sub> = 34Hz), i.e. downfield [Slack et al. *Inorg Chem* **18** 3125 1979]. Its absolute crystal X-ray structure has been determined and is the one predicted, i.e. the chelate ring is λ, the methyl group is equatorially disposed and the absolute configuration of the diphosphine is *R*. The configuration in solution is assumed to be the same, being consistent with the stereochemistry of the catalytically produced products.

Like the above rhodium complexes of *S,S*-chiraphos, this *R*-prophos rhodium complex is an efficient homogeneous catalyst for the production of α-amino acids, and the optical yields appear to be insensitive to the nature of the substituents on the substrates which provide the “natural” *S*-amino acids in 90±3% optical yields and high chemical yields (<87%) [compare with the *S,S*-chiraphos rhodium precatalysts above which gives the non-natural *R*-α-amino acids]. Furthermore, the catalyst *R*-prophos rhodium complex can breed its own chirality so that large quantities of *R*-prophos can be made from the catalytic hydrogenation of ethyl acetoxyacrylate, *via* ethyl *S*-(-)-O-acetyllactate, by the *R*-prophos catalyst itself. This procedure has been used to produce *S*-prophos

if the pre-catalyst used is  $[Rh(S,S\text{-chiraphos})(NBD)]ClO_4$ . [Fryzuk & Bosnich *J Am Chem Soc* **100** 5491 1978.] An ingenious application of this *R*-prophos rhodium pre-catalyst using hydrogen, deuterium and tritium has been adopted for the asymmetric synthesis of chiral lactic acid in which the methyl group is chiral by virtue of its having a hydrogen, a deuterium and a tritium atom on the methyl carbon atom. The orientations here being opposite to those obtained with the *S,S*-chiraphos rhodium precatalysts. [Fryzuk & Bosnich *J Am Chem Soc* **101** 3043 1979.]

Note that these rhodium catalysts are in fact *pre-catalysts* because in the presence of hydrogen their {e.g.  $[Rh(S,S\text{-chiraphos})(NBD)]^+$ } strong deep red-orange colours are reduced to a light straw yellow colour of the true catalyst {e.g.  $[Rh(S,S\text{-chiraphos})(H)_2]^+$ }. The mechanism of these asymmetric hydrogenations has been studied in some detail [Brown & Chaloner *J Am Chem Soc* **102** 3040 1980, Brown et al. *J Organomet Chem* **216** 263 1981].

**Rhodium(III) chloride**  $[RhCl_3 \pm x H_2O]$  [10049-07-7; 20765-98-4  $x H_2O$ ] **M 209.3 (anhydrous), m 100° (hydrate, dec)**. The anhydrous form is crystalline and is hydrophobic. It is a red powder insoluble in  $H_2O$ , but soluble in aqueous NaOH or cyanide solutions. The hydrate, however, is soluble in  $H_2O$ . [cf preparation of the trihydrate by Anderson & Basolo *Inorg Chem* **7** 214 1963.]

It catalyses the direct conversion of methane to AcOH [Sen et al. *Nature* **368** 613 1994], and is a catalyst for conjugate reduction of cinnamaldehydes followed by cross-coupling with arylboronic acids [Wang et al. *JCS, Chem Commun* 1192 2004]. It also catalyses the deuteration of saturated hydrocarbons [Takahashi et al. *J Am Chem Soc* **97** 7489 1975], and in the presence of  $Ph_3P$  (Wilkinson's catalyst) the trimethylsilylation of methacrylic esters to 1-alkoxy-1-trimethylsilyl oxyethylenes with rearrangement of a methyl group in the presence of  $Me_3SiH$  in THF is achieved [see Revis & Hilty *J Org Chem* **55** 2972 1990.]

**Rhodium(II) pentafluorobenzoate dimer [tetrakis(pentafluorobenzoato)dirhodium(II)]** [75863-37-5] **M 947.2**. This dimer is obtained from a solution of perfluorobenzoic acid (4.5g) and hydrated  $RhCl_3$  (1g) in EtOH (80ml), to which is added NaOH (0.8g), and heated under reflux for 3 hours, cooled and the solid is filtered off. The solid is refluxed for a further hour with fresh EtOH (80ml) and is filtered again. The combined EtOH solutions are evaporated *in vacuo*, the solid green residue is extracted with several volumes of  $Et_2O$  until the extracts are colourless; the  $Et_2O$  is concentrated and chromatographed on  $SiO_2$  (eluting with toluene/ $Et_2O$  9:1). Some of the complex (1g) is collected in one of the fractions, and dried *in vacuo* for 3 hours [ $150^\circ$  at  $10^{-2}$  mm] and is analytically pure. The IR(KBr) has  $\nu_{max}$  at 1655m, 1597s, 1525m, 1500s, 1433s, 1405s, 1297w, 1118m, 997s, 942w, 768m  $cm^{-1}$ . Like the triflate salt (following entry) it catalyses the addition of carbenes (drived from alkyl diazo diazoesters) to aromatic compounds, although its catalytic efficiency is unaffected by the bulk of the diazo ester alkoxy group [Anciaux et al. *J Org Chem* **46** 873 1981].

**Rhodium(II) trifluoroacetate dimer [dirhodium tetra-trifluoroacetate, tetrakis(trifluoroacetato)dirhodium(II),  $Rh_2(Tf)_4$ ]** [31126-95-1] **M 657.9**. It can be purified in much the same way as the preceding acetate dimer (see [15956-28-2]) and is recrystallised from  $*C_6H_6$  before use. It is a homogeneous catalyst [Black *Aldrichimica Acta* **15** 13 1982]. Tetrakis(perfluoroacetato)dirhodium(II) catalyses the addition of carbenes (generated from diazo esters) to aromatic compounds at room temperature to generate 1-carbalkoxycyclohepta-2,4,6-trienes (with high kinetic selectivity for the *nonconjugated* isomers) in very good yields. A competitive reduction of the catalyst occurs simultaneously causing the reaction to slow down, and should be taken onto account [Anciaux et al. *J Org Chem* **46** 873 1981].

**Ruthenium(III) acetylacetonate**  $[Ru(acac)_3]$  [14284-93-6] **M 398.4, m 240°(dec)**. Purify the complex by recrystallisation from  $*benzene$ . [Wilkinson *J Am Chem Soc* **74** 6146 1952, *Beilstein* **1** IV 3677.] It catalyses the hydrogenation of dimethyl oxalate to ethylene glycol under mild conditions [70bar  $H_2$  pressure at  $100^\circ$ ] and the best ligand is  $MeC(CH_2PPh_2)_3$  with a trace of Zn (0.07% of oxalate) in MeOH with yields of >84% and with turnover numbers of 875 (turnover frequency/hour of 53.5) [Teunissen & Elsevier *J Chem Soc, Chem Commun* 667 1997].

**Ruthenium (benzylidene)dichloro-bis-(tricyclohexylphosphine) [phenylmethylene-bis-(tricyclohexylphosphine) dichlororuthenium (Grubbs catalyst—first generation)]** [172222-30-9] **M 823.0, m 153° (dec)**. Wash

it repeatedly with Me<sub>2</sub>CO and MeOH and dry it in a vacuum. *Alternatively*, dissolve it in warm CH<sub>2</sub>Cl<sub>2</sub>, concentrate it to half its volume, filter, add MeOH to precipitate it as purple microcrystals. Filter these off, wash several times with Me<sub>2</sub>CO and MeOH and dry them in a vacuum for several hours. [Schwab et al. *J Am Chem Soc* **118** 100 1996, Miller et al. *J Am Chem Soc* **118** 9606 1996, Furstner & Langermann *J Am Chem Soc* **119** 9130 1997.] It is used to catalyse ring-closing metathesis [Schrodi & Pederson *Aldrichimica Acta* **40** 45 2007, Schmidt *Angew Chem. Int Ed* **42** 4997 2003], and promotes olefin metathesis with ruthenium based catalysts [Grubbs *Tetrahedron* **60** 7117 2004]

§ A polymer supported version is available [Schwab et al. *Angew Chem, Intl Edn* **34** 2039 1995].

**Ruthenium(III) chloride (RuCl<sub>3</sub>)** [*3H<sub>2</sub>O* 13815-94-6; *xH<sub>2</sub>O* 14898-67-0; *Anhydrous* 10049-08-8] **M 207.4 (anhydrous), 261.5 (3H<sub>2</sub>O), d<sup>20</sup> 3.11.** The anhydrous salt exists in two forms. The α-form is produced by the slow reaction of Cl<sub>2</sub> with Ru metal in siliceous containers at >600° to give black lustrous hexagonal crystals which are antiferromagnetic and insoluble in H<sub>2</sub>O or EtOH [Hill & Beamish *J Am Chem Soc* **72** 4855 1950]. The second β-form is prepared by heating Ru metal in a stream of CO and Cl<sub>2</sub> at 340°, to avoid the formation of carbonyl compounds such as Ru(CO)<sub>2</sub>Cl<sub>2</sub> and is free from the metal or α-RuCl<sub>3</sub>. It is formed in dark brown fluffy hexagonal crystals that are soluble in EtOH. It is the metastable form because at or about the transition temperature of 450° the β-form is slowly converted to the α-form (irreversibly, with *t*<sub>0.5</sub> ~1 hour), and is the best way to prepare the α-form. [Fletcher et al. *Nature* **199** α-form 1089 1963.]

**Hydrated RuCl<sub>3</sub>** is one of the most useful inorganic ruthenium compounds, and particularly for the preparation of Ru coordinated compounds. It is prepared by evaporating RuO<sub>4</sub> in concentrated hydrochloric acid in a stream of HCl gas. Unlike the anhydrous form, the hydrate is soluble in H<sub>2</sub>O, but a fresh aqueous solution (brown to brown-green colour) does not precipitate AgCl with AgNO<sub>3</sub> solution because the halogen atoms (as well as one molecule of H<sub>2</sub>O) are coordinated to the metal. However, on warming in H<sub>2</sub>O the halogen atoms are displaced by H<sub>2</sub>O molecules and the Cl<sup>-</sup> ions can be precipitated (or titrated) with AgNO<sub>3</sub> [Grube & Fromm *Z Electrochem* **45** 661 1940, cf Connick & Fine *J Am Chem Soc* **83** 3414 1961].

It is a soluble catalyst used for the oxidative cyclisation of 1,7-dienes to oxepane diols [Piccialli et al. *Tetrahedron Lett* **48** 5131 2007], and promotes a site-specific hydroxylation of tertiary carbon-hydrogen bonds of cyclic ethers in the presence of periodate or bromate [Lee et al. *J Org Chem* **72** 5820 2007].

It also catalyses oxidation reactions with an oxidant, e.g. the methylene group in cyclopropylmethyl-compounds to a carbonyl group in the presence of metaperiodate [Hasegawa et al. *Chem Lett (Jpn)* 1385 1985, Carlsen et al. *J Org Chem* **46** 3936 1981, cf Review Gore *Platinum Metals Rev* **27** 111 1983]. It catalyses the synthesis of 2-ethyl-3-methylquinolines from primary aromatic amines and triallylamine [Cho et al. *Tetrahedron Lett* **40** 1499 1999], and has been used for the selective hydrogenation of unsaturated aldehydes [Fujita et al. *J Catal* **255** 95 2004].

**Ruthenium [1,3-bis-(2,4,6-trimethylphenyl)-2-imidazolidinylidene] dichloro (benzylidene) (tri-cyclohexylphosphine) {Benzylidene-[1,3-bis(2,4,6-trimethylphenyl)-4,5-dihydro-2-imidazolinyldene]dichloro-(tri-cyclohexylphosphine)-ruthenium, [(1,3-Bis(2,4,6-trimethylphenyl)-2-imidazolinyldene) dichloro (phenyl-methylene)(tricyclophosphine)ruthenium, (Grubbs catalyst—second generation)}** [246047-72-3] **M 849.0, m 143.5-148.5°.** This catalyst is prepared by adding a solution of *tert*-BuOK (1.25g, 10.2mmol, 1.4 equiv) in dry THF (200ml) slowly to a suspension of 1,3-dimesityl-4,5-dihydroimidazolium tetrafluoroborate (4.2g, 10.2mmol, 1.4 equivalent) in dry THF (100ml) in a flame dried Schlenk flask (500ml) under N<sub>2</sub> (magnetic stirring) at ~20°. The BF<sub>4</sub><sup>-</sup> salt dissolves immediately, the cloudy yellow mixture is stirred at ~20° for 1 hour, and then transferred (cannula) to a larger Schlenk flask (twice size) under argon. To this is added RuCl<sub>2</sub>(=CHPh)(PCy<sub>3</sub>)<sub>2</sub> (7.29, 7.29mmol, 1.0 equivalents) in dry \*C<sub>6</sub>H<sub>6</sub> (400ml) and heated at 80° for 30 minutes when the reaction is complete (by <sup>1</sup>H NMR). The volatiles are removed *in vacuo* and the residue is washed with dry MeOH, or pentane (4 × 100ml) to provide the active catalyst as a pink-brown microcrystalline powder (4.64g, 75%). It has <sup>1</sup>H NMR (400MHz, CD<sub>2</sub>Cl<sub>2</sub>) with δ at 19.16 (s, 1H), 7.37-7.05 (m, 9H), 3.88 (s, 4H), 2.56—0.15 (m, 51H); <sup>31</sup>P NMR (161.9MHz, CD<sub>2</sub>Cl<sub>2</sub>) with δ at 31.41 and HRMS (FAB) at 848.3306 for M<sup>+</sup> [Scholl et al. *Org Lett* **1** 953 1999].

It is used in ruthenium catalysed ring closure metathesis (RCM) and olefin metathesis [Kulkarni & Diver *J Am Chem Soc* **126** 8110 2004, Schmidt *Angew Chem. Int Ed* **42** 4996 2003, Scholl et al. *Org Lett* **1** 953 1999], and generate unsaturated sultones (e.g. 2,7-3*H*-dihydro-[1,2]oxathiepine-2,2-dioxide) from olefinic sulfonates (e.g. but-3-enyl allylsulfonate) *via* ring closure metathesis [Le Flohic et al. *Tetrahedron* **62** 9017 2006].

**Ruthenium [(R-P-Phos)(acac)<sub>2</sub>] {Ruthenium [R-(+)-2,2',6,6'-tetramethoxy-4,4'-bis(diphenylphosphino)-3,3'-bipyridine)-bis(acetylacetonate)]** is prepared by mixing equimolar amounts of R-P-Phos (see R-(+)-[221012-82-4] in Part-2 below) with Ru(acac)<sub>3</sub> [14284-93-6] in the presence of a reducing agent (Zn powder) in refluxing EtOH to give the solid catalyst in 97% yield which was characterized by <sup>1</sup>H, <sup>31</sup>P NMR, elemental analysis and X-ray crystallographic structure determination. **Ru[(S-P-Phos)Cl<sub>2</sub>](DMF)<sub>n</sub>** is prepared in a similar manner with the appropriate ingredients (using S-(-)-P-Phos [362524-23-0]) in DMF. [Pai et al. *J Am Chem Soc* **122** 11513 2000.] These (see also below) atropisomeric bi-heteroaromatic diphosphine (P-Phos and its variants) complexes with transition metals are effective in catalysing asymmetric reactions including asymmetric hydrogenation of 2-arylacrylates, β-ketoesters, arylketones, hydrosilylation and C-C bond formation with high stereospecificity. [Wu & Chan *Acc Chem Res* **39** 711 2006, Au-Yeung & Chan *Coordination Chemistry Reviews* **248** 2151 2004, Pai et al. *J Am Chem Soc* **122** 11513 2000.]

**Ruthenium [(R-Xylyl-P-Phos)(C<sub>6</sub>H<sub>6</sub>)Cl<sub>2</sub>] {Ruthenium R-(+)-2,2',6,6'-tetramethoxy-4,4'-bis[di(3,5-xylyl)phosphino]-3,3'-bipyridine)(benzene) dichloride]** is prepared in 88% yield by mixing equimolar amounts of [RuCl<sub>2</sub>(C<sub>6</sub>H<sub>6</sub>)<sub>2</sub>] [37366-09-9] with R-Xylyl-P-Phos (see R-(+)- [442905-33-1] in Part-2 below) in EtOH/\*C<sub>6</sub>H<sub>6</sub> (8:1) and heating at 50-60° for 1 hour. This ruthenium complex is a highly active, enantioselective, air stable catalyst for the asymmetric hydrogenation of β-ketoesters (PH<sub>2</sub> = 200-350psi at 70-90° for 1-15 hours) to optically active β-hydroxyesters with ~100% conversion and >91% enantiomeric enrichment. It is very stable in the solid state in air; and a stirred solution for 10 hours under air showed no loss of activity or stereoselectivity; and its <sup>31</sup>P NMR (200MHz, CDCl<sub>3</sub>) is not different from that obtained for the original sample under N<sub>2</sub> which has an AB set of signals with δ<sub>p</sub> at 33.49 (d, J<sub>PA-PB</sub> = 62.9Hz) and 39.96 (d, J<sub>PA-PB</sub> = 62.5Hz). The coordination of ruthenium to benzene makes the phosphorus atoms non-equivalent. The resonance observed in the <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) at 5.65ppm is assignable to the 6H resonances of the η<sup>6</sup>-coordinated benzene moiety [Wu et al. *Tetrahedron Lett* **43** 1539 2002].

Other P-Phos related ligands, such as the chiral 4-bis(tolylphosphino) (**Tol-P-Phos**) and chiral 4-bis(cyclohexylphosphino) (**Cy-P-Phos**) ligands, were prepared and purified in a similar, or slightly modified, manner using the appropriate phosphines. They also form ruthenium complexes that have related catalytic activities [Wu & Chan *Acc Chem Res* **39** 711 2006, Au-Yeung & Chan *Coordination Chemistry Reviews* **248** 2151 2004].

**Ruthenium(II)-tris(triphenylphosphine) dichloride [tris(triphenylphosphine)ruthenium(II) dichloride] [15529-49-4] M 958.8; 433 (in Me<sub>2</sub>CO/N<sub>2</sub>), m 132-134°.** This catalyst is prepared by heating under reflux RuCl<sub>3</sub>·3H<sub>2</sub>O (0.26g, 1mmol) and Ph<sub>3</sub>P (1.57g, 6mmol) in MeOH (65ml) under N<sub>2</sub> for 3 hours, when the reddish-brown tris-complex separates. It is washed well with MeOH, then Et<sub>2</sub>O and dried *in vacuo* at 60° (yield 75%). It is soluble in CHCl<sub>3</sub>, Me<sub>2</sub>CO, \*C<sub>6</sub>H<sub>6</sub>, EtOAc and hot *i*-PrOH, but insoluble in H<sub>2</sub>O, Et<sub>2</sub>O and sparingly soluble in alcohol. When less Ph<sub>3</sub>P is used in e.g. MeOH, complexes such as Ru(III)-di(Ph<sub>3</sub>P)(MeOH)Cl<sub>2</sub> are formed. *Note that if the above mixture is shaken, not refluxed, Ru(II)-tetrakis(triphenylphosphine) dichloride is formed which turns green on exposure to air.* [Stephenson & Wilkinson *J Inorg Nucl Chem* **28** 945 1966, Sammakia & Stangeland *J Org Chem* **62** 6105 (footnote) 1997.]

The *tris-complex* catalyses a variety of reactions including hydrogen transfer [Sasson & Blum *Tetrahedron Lett* 2167 1971, Regen & Whitesides *J Org Chem* **37** 1832 1972], insertion of *t*-BuOO- into the carbon atom α to the N-atom of, e.g. *N*-methoxycarbonylpyrrolidine and 1,2,3,4-tetrahydroisoquinoline with *t*-BuOOH [Kondo et al. *J Org Chem* **56** 487 1991], intramolecular cyclisation of unsaturated α,α-dichloroesters or amides [Hayes et al. *J Org Chem* **51** 5501 1986], and the reduction of CHO by HCOOH/Et<sub>3</sub>N/THF without affecting NO<sub>2</sub>, RCOR', ester or *tert*-amide groups [Kai & Arcelli *Tetrahedron Lett* **26** 3365 1985]; and for other examples of catalysis by this Ru-complex see selected volumes of Fieser & Feiser's *Reagents in Organic Synthesis*, and *Angew Chem. Int Ed* **46** 1905 2007. When complexed further with chiral (phosphinoferrocenyl)oxazolines, hydrogen transfer is not only considerably enhanced but is also highly stereospecific [see further in this chapter].

**Samarium (II) iodide (SmI<sub>2</sub>)** **Samarium (II) iodide (SmI<sub>2</sub>)** see [32248-43-4], Chapter 5, Inorganic Compounds.

**Silver acetate (AgOAc)** see [563-63-3], Chapter 5, Inorganic Compounds.

**Silver tetrafluoroborate** [AgBF<sub>4</sub>] see [14104-20-2], Chapter 5, Inorganic Compounds.

**Tetrakis(triphenylphosphino)palladium(0)** [**palladium-tetrakis(triphenylphosphine)**] [14221-01-3] **M 1155.6, m 100-105°, 115°, ~116° (dec, sealed tube under N<sub>2</sub>), (m is unreliable and is not a criterion of purity because it varies).** This catalyst is prepared from PdCl<sub>2</sub> (5.9g, 0.1mol), Ph<sub>3</sub>P (43.7g, 0.5ml) in Me<sub>2</sub>SO (400ml) under a vacuum N<sub>2</sub> system with pressure release, and the yellow mixture is heated (oil bath) with stirring until dissolution (140°). The bath is removed, the mixture is stirred rapidly for 15 minutes and hydrazine hydrate (6.7g 0.4mol) is injected within 1 minute. A vigorous evolution of N<sub>2</sub> occurs and the dark solution is rapidly cooled with a cool water bath and crystallisation takes place (~125°). When the temperature reaches ~20° the solid is filtered off (coarse sintered glass) under N<sub>2</sub>, washed with EtOH (2 × 50ml), Et<sub>2</sub>O (2 × 50ml) and the yellow crystalline solid is dried by a slow stream of N<sub>2</sub> through the funnel overnight (37.4g 97%). *Alternatively*, it can be prepared freshly by mixing Pd(NO<sub>3</sub>)<sub>2</sub> (2mmols) and PPh<sub>3</sub>(2mmols) in hot \*C<sub>6</sub>H<sub>6</sub> when vigorous evolution of nitric oxide occurs (fume cupboard) and a solid mass separates. This is collected and crystallised from EtOH. It should not be heated excessively as it dissociates to Pd(PPh<sub>3</sub>)<sub>3</sub> and PPh<sub>3</sub>, and then further to Pd(PPh<sub>3</sub>)<sub>2</sub> and PPh<sub>3</sub>. It is also air sensitive as PPh<sub>3</sub> is oxidised to PPh<sub>3</sub>O. It is stable only for short periods because on exposure to heat or air it turns from yellow to orange and dissociates in solution, so the solutions should be used directly. Its cryoscopic constant in \*C<sub>6</sub>H<sub>6</sub> (at 0.601g/20ml) corresponds to **M 1156** [Malatesta & Angoletti *J Chem Soc* 1186 1957]. Molecular weight determination in \*C<sub>6</sub>H<sub>6</sub> indicates considerable dissociation, and the solution absorbs O<sub>2</sub> rapidly to give an insoluble green oxygen complex [Nyman et al. *J Chem Soc A* 561 1968]. It is moderately soluble in \*C<sub>6</sub>H<sub>6</sub> (5g/100ml), CH<sub>2</sub>Cl<sub>2</sub> and CHCl<sub>3</sub> but less soluble in Me<sub>2</sub>CO, THF and MeCN. Pd(Ph<sub>3</sub>P)<sub>4</sub> may be handled in air but it is best stored under N<sub>2</sub>. [Coulson *Inorg Synth* 28 121 1990, Malatesta & Angoletta *J Chem Soc* 1186 1990, *Beilstein* 16 IV 954.]

Pd(Ph<sub>3</sub>P)<sub>4</sub> is a very versatile catalyst for promoting the dimerisation of butadiene to 1,3,7-octatriene [Takahashi et al. *Bull Chem Soc Jpn* 41 454 1968], catalysing various coupling reactions, without homo-coupling occurring [Brocato et al. *Tetrahedron Lett* 33 7433 1992, Arcadi et al. *Tetrahedron Lett* 34 2813 1993, McClure & Danishefsky *J Am Chem Soc* 115 6093 1993, Paquette & Astles *J Org Chem* 58 165 1993, Schoo et al. *Org Lett* 8 4141 2006], including Suzuki coupling [Trost *Tetrahedron* 33 2615 1977]. [*Beilstein* 16 IV 954.]

This palladium catalyst bound to a polymer support (~0.06mmol/g) is also commercially available [*cf* Frnger & LeDrain *Tetrahedron Lett* 39 4287 2988].

**Tetra-*n*-propylammonium perruthenate (TPAP, tetrapropyl tetraoxoruthenate)** [114615-82-6] **M 351.4, m 160°(dec).** This stable dark green solid is a useful catalyst for a variety of oxidation reactions at about room temperature. It is used at a concentration of about 5 mol% with an equivalent of *N*-methylmorpholine *N*-oxide which provides the required oxygen with typical turnovers of about 250. The reactions are carried out in CH<sub>2</sub>Cl<sub>2</sub> and/or MeCN, although the latter assists the reactions in CH<sub>2</sub>Cl<sub>2</sub> considerably. Care should be taken to add TPAP slowly as the reaction can be vigorous. Addition of finely ground 4Å molecular sieves moderate the rate and efficiency of the reaction. The reactions usually proceed to completion within 5 minutes to one hour. Cooling and slow addition of TPAP is necessary with reactions on a large scale. Workup is simple, and is carried out by passing the solution through a short silica-gel column and eluting with EtOAc followed by evaporation and crystallisation or distillation of the product. When MeCN is the solvent, or is present in the solvent, it is advisable to evaporate the solvent prior to passage through the silica-gel column. [Ley et al. *Synthesis* 639 1994; see entry in the “Metal-Organic Compounds” chapter for further information].

**Triphenylphosphinegold(I) bis(trifluoromethanesulfonyl)imidate** ([Ph<sub>3</sub>PAu(I)]<sup>+</sup> [NTf<sub>2</sub>]<sup>-</sup>) [866395-16-6] **M 739.4, m >230°.** AgNTf<sub>2</sub> is first prepared by mixing Ag<sub>2</sub>CO<sub>3</sub> (276mg, 1.0mmol) and HNTf<sub>2</sub> (562mg, 2.0mmol, 2 equivalents [82113-65-3]) in H<sub>2</sub>O (10ml), and refluxing for 3 hours (to eliminate CO<sub>2</sub>); then evaporating to dryness *in vacuo*. It can be used as a solid or dissolved in dry CH<sub>2</sub>Cl<sub>2</sub> (5ml). [Li et al. *J Org Chem* 73 4323 2008.] The versatile catalyst is then prepared by dissolving Ph<sub>3</sub>PAuCl (2mmol, [14243-64-2]) in CH<sub>2</sub>Cl<sub>2</sub> (5ml) and adding to it solid AgNTf<sub>2</sub> (786mg, 2mmol) or the preceding CH<sub>2</sub>Cl<sub>2</sub> (5ml) solution. Immediate separation of AgCl occurs. After stirring for 15 minutes the <sup>31</sup>PMR reveals a single peak of the Au salt. The AgCl is removed by filtration through Celite, the pale coloured solution is evaporated *in vacuo*, and the residue is dried *in vacuo* to give a quantitative yield of the catalyst. It forms small white crystals which are air stable, but should be preferably stored in a dry inert atmosphere. It is soluble in Me<sub>2</sub>CO, CH<sub>2</sub>Cl<sub>2</sub>, and CHCl<sub>3</sub> but insoluble in petroleum ethers. Its IR (CCl<sub>4</sub>) has ν<sub>max</sub> at 1482, 1437, 1405, 1384, 1216, 1196, 1133, 1103 and 960, cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) has δ<sub>H</sub> at 7.47-7.55 (m, 12H, aromatic H) and 7.56-7.62 (m, 3H, aromatic H close to

Au); the  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 118.4 (q,  $J_{^{13}\text{C}-^{19}\text{F}} = 323.0\text{Hz}$ ), 126.3 (d,  $J_{^{13}\text{C}-^{31}\text{P}} = 65.8\text{Hz}$ ), 128.5 (d,  $J_{^{13}\text{C}-^{31}\text{P}} = 12.5\text{Hz}$ ), 133.5 (d,  $J_{^{13}\text{C}-^{31}\text{P}} = 2.5\text{Hz}$ ) and 133.1 (d,  $J_{^{13}\text{C}-^{31}\text{P}} = 13.7\text{Hz}$ ); and the  $^{31}\text{P}$  NMR (121.5MHz,  $\text{CD}_2\text{Cl}_2$ ) has  $\delta_{\text{P}}$  at 30.7. [Mésailles et al. *Org Lett* **7** 4133 2005.]  $\text{NTf}_2$  is a weakly coordinating counter ion which confers stability to the complex.  $^{31}\text{P}$  NMR spectra of pre-formed solutions of the complexes  $\text{Ph}_3\text{PAuBF}_4$ ,  $\text{Ph}_3\text{PAuPF}_6$  and  $\text{Ph}_3\text{PAuSbF}_6$ , using the respective Ag salts, showed that they were not very stable. Attempts to isolate the complexes were unsuccessful. [Mésailles et al. *Org Lett* **7** 4133 2005.]

At concentrations of 1-2mol% in  $\text{CH}_2\text{Cl}_2$  or  $\text{Me}_2\text{CO}$ ,  $\text{Ph}_3\text{PAu(I)}^+ [\text{NTf}_2]^-$  catalyses the stereoselective isomerisation of butynediol monobenzoates into functionalised 2,5-dihydrofurans (in a sequence of two steps) in high yields at room temperature and in 15 to ~30 minutes [Buzas et al. *Org Lett* **8** 1957 2006]. It is exceedingly active in catalysing a wide variety of *ene* cycloisomerisations [Mésailles et al. *Org Lett* **7** 4133 2005]. Also under similar mild conditions, in the presence of electrophilic bromine (from *N*-bromosuccinimide) or iodine (from *N*-iodosuccinimide), it catalyses the formation of linear  $\alpha$ -halo-enones from propargyl acetates [Yu et al. *Tetrahedron* **65** 1846 2009], and 4-haloalkylidene-1,3-dioxalane-2-ones with propargyl *tert*-butylcarbonates [Buzas et al. *Tetrahedron* **65** 1889 2009], products which are suitable substrates for Pd-catalysed cross coupling reactions. Like the following  $\text{OTf}^- \text{Au}$  catalyst, it promotes the glycosylation of 1,2-anhydrosugars as donors using protected sugars with one free sterically unhindered OH group [Li et al. *J Org Chem* **73** 4323 2008].

**Triphenylphosphinegold(I) trifluoromethanesulfonate** ( $[\text{Ph}_3\text{PAu(I)}]^+ [\text{OTf}]^-$ ) [156397-47-6] **M 608.2, m >230°**.  $\text{Ph}_3\text{PAu(I)}^+ [\text{OTf}]^-$  differs from the preceding catalyst in the counter anion, and is much less air and moisture sensitive. Due precautions need to be taken in use and storage. It is prepared by dissolving  $\text{Ph}_3\text{PAuCl}$  (0.1mmol, [14243-64-2]) and silver trifluoromethanesulfonate (0.1mol, [2923-28-6]) in dry  $\text{CH}_2\text{Cl}_2$  (1ml) and stirred for 5 minutes. Filtration from AgCl provides a 0.1M solution of catalyst in  $\text{CH}_2\text{Cl}_2$ . Generally the AgCl does not interfere in the catalytic process and so the catalyst can be prepared *in situ*. Colourless  $\text{Ph}_3\text{PAu(I)}^+ [\text{OTf}]^-$  can be obtained by evaporating the  $\text{CH}_2\text{Cl}_2$  solution and stored appropriately. It is a superior catalyst than anhydrous  $\text{ZnCl}_2$  for the well established glycosylation reactions with 1,2-anhydrosugars as donors [Li et al. *J Org Chem* **73** 4323 2008], and it catalyses the intramolecular hydroamination of terminal alkenes in high yields with 1-5mol% of catalyst by heating in toluene at 100° for 12-48 hours, or the intra- and inter- molecular hydroamination in  $\text{ClCH}_2\text{CH}_2\text{Cl}$  by microwave radiation as heat source in ~30 minutes in high yield [Liu et al. *Org Lett* **8** 2707 2006]. In a useful application, the catalyst  $\text{Ph}_3\text{PAu(I)}^+ [\text{OTf}]^-$  promotes a cascade cyclisation/oxidative cleavage of a carbon-carbon triple bond in *Z*-enynols, e.g. 5-Ph or 5-*n*-butyl-pent-2-ene-4-yne-1-ols, in the presence of molecular oxygen to give high yields of the corresponding 2,5-dihydrofuran-2-ones and releasing C-5 with its substituent as the aldehyde or acid. The reaction involves free radicals as it is completely suppressed in the presence of the radical scavenger 4-hydroxy-TEMPO [Liu et al. *J Am Chem Soc* **128** 11332 2006].

**Tris(dibenzylideneacetone)dipalladium(0) chloroform adduct**  $[\text{Pd}_2(\text{dba})_3(\text{CDCl}_3)]$  [52522-40-4] **M 1035.1, m 122-124°; solvent free  $\text{Pd}_2(\text{dba})_3$**  [51364-51-3] **M 95.7, m 152-155°**. When  $\text{PdCl}_2$  (1.05g, 5.92mmol) is added to a solution of dibenzylidene acetone (dba, 4.60g, 19.6mmol) and NaOAc (3.9g, 47.5mmol) in MeOH (150ml at 50°) and mixed at 40° for 4 hours then cooled, a reddish-purple precipitate separates. This is collected, washed with  $\text{H}_2\text{O}$  then  $\text{Me}_2\text{CO}$  and dried *in vacuo*. The precipitate (3.39g) is purified by dissolving it in hot  $\text{CDCl}_3$  (120ml), filtering, and to the deep violet filtrate is added slowly  $\text{Et}_2\text{O}$  (170ml) when the *chloroform adduct* separates as deep purple needles which are collected and dried *in vacuo* (80% yield, m 122-124°). Recrystallisation from  $^*\text{C}_6\text{H}_6$  instead of  $\text{CHCl}_3$  gives deep-violet needles of **tris(dibenzylideneacetone)-dipalladium(0) benzene adduct** (62.5% yield, m 141-142°). By using toluene instead of  $\text{CHCl}_3$ , deep-violet needles of **tris(dibenzylideneacetone) dipalladium(0) toluene adduct** (36% yield, m 140-141°) are obtained. The solvates can be exchanged with each other (i.e.  $\text{CDCl}_3$ ,  $^*\text{C}_6\text{H}_6$  or toluene) without affecting the coordinating reactivity of the complex with other ligands such as  $\text{Ph}_3\text{P}$ , Bipy, *o*-phen, olefins, *o*-quinones etc. [Ukai et al. *J Organomet Chem* **65** 253 1974, for applications see Fustero *Org Lett* **8** 4129 2006.]

If the procedure is used with bis(*p*-methoxybenzylidene)acetone and the complex crystallised from  $\text{CDCl}_3/\text{Et}_2\text{O}$ ,  **$\text{Pd}_2(p,p'$ -methoxy-dba) $_3$**  is obtained as deep violet needles (72.3% yield, m 141-143°) which do not complex with the solvent. [cf  **$\text{Pd}(\text{dba})_2$**  above.] [Ukai et al. *J Organomet Chem* **65** 253 1974.]

**Tris(dibenzylideneacetone)palladium**  $[\text{Pd}(\text{dba})_3]$ : When a fourfold excess of dba is added to solution of  $\text{Pd}(\text{dba})_2$  or  $\text{Pd}_2(\text{dba})_3$  in  $^*\text{C}_6\text{H}_6$ , heated and reduced in volume, the dark red colour of  $\text{Pd}(\text{dba})_2$  turns to brown,

and on complete evaporation, orange brown crystals of **Pd(dba)<sub>3</sub>** contaminated with yellow crystals of dba are obtained. Washing the crystals with \*C<sub>6</sub>H<sub>6</sub> leads to crystal decomposition. The IR of the complex has  $\nu_{\max}$  at 1651 (C=O) and olefine bands at 1598, 1580 and 1531 cm<sup>-1</sup>. X-ray analysis of the Pd(dba)<sub>3</sub>—\*C<sub>6</sub>H<sub>6</sub> crystals showed that each pentadienone ligand is coordinated through one olefin group, and the Pd atom is trigonal with  $\sim C_3$  symmetry [Mazza & Pierpont *Inorg Chem* **12** 2955 1973, cf *Handbook of Organopalladium for Organic Synthesis* Negishi ed. Wiley, Hoboken NJ 2002, ISBN 0-471-31506-0].

**Tris[triphenylphosphinegold(I)]oxonium tetrafluoroborate ([Ph<sub>3</sub>PAu(I)]<sub>3</sub>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup>)** [53317-87-6] **M 1480.6, m 207° (dec), 207-208° (dec), 220-221° (dec)**. The catalytic oxonium salt is prepared by adding a solution of AgBF<sub>4</sub> (0.2g, 1.03mmol) in MeOH (5ml) to a solution of Ph<sub>3</sub>PAuCl (0.5g, 1.01mmol, [14243-64-2]) in THF (20ml) and the AgCl that precipitated is filtered off. A solution of KOH (0.1g, 1.78mmol) and NaBF<sub>4</sub> (0.5g, 4.55mmol) in MeOH (100ml) is added to the filtrate, stirred for 1 hour, the solvent is evaporated off *in vacuo*, the residue is extracted with CHCl<sub>3</sub> (2 x 30ml), the combined extract is filtered, and hexane (100ml) is added to precipitate the oxonium salt which gives analytically pure [Ph<sub>3</sub>PAu(I)]<sub>3</sub>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup> (0.42g, 84%), **m 220-221° (dec)**, after recrystallisation from a saturated CHCl<sub>3</sub> solution on adding  $\sim 1.5$ -fold volume of Me<sub>2</sub>CO. **Alternatively**, Ag<sub>2</sub>O freshly prepared from AgNO<sub>3</sub> (5.0g, 29.5mmol), and NaBF<sub>4</sub> (5.0g, 45.5mmol, finely powdered) are added to a solution of Ph<sub>3</sub>PAuCl (4.0g, 8.08mmol) in Me<sub>2</sub>CO (600ml), stirred vigorously for 1 hour, and the solvent evaporated *in vacuo*. The residue is extracted with \*C<sub>6</sub>H<sub>6</sub> to removed unreacted Ph<sub>3</sub>PAuCl, and the oxonium salt is extracted with CHCl<sub>3</sub> (3 x 40ml), filtered, and hexane (400ml) is added to crystallise [Ph<sub>3</sub>PAu(I)]<sub>3</sub>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup> (3.6g, 90%) out. It is soluble in CHCl<sub>3</sub> and CH<sub>2</sub>Cl<sub>2</sub> but insoluble in hexane and Et<sub>2</sub>O. Its IR (nujol mull) has  $\nu_{\max}$  at 1050-1070 (br, BF<sub>4</sub><sup>-</sup>) cm<sup>-1</sup>. It is a versatile air and moisture tolerant catalyst that has been used in a variety of reactions (see below). X-ray crystallography showed that in the crystals the oxonium ions are dimeric, with the two pyramidal monomeric (Ph<sub>3</sub>PAu)<sub>3</sub>( $\mu_3$ -O)<sup>+</sup> fragments interacting *via* Au—Au' bonds ( $\sim 3.16\text{\AA}$ ) involving two of the three Au atoms in each unit. The pyramidal structure has the O' atom centrally above the Au<sub>3</sub> triangular plane, and the resulting six-membered Au<sub>2</sub>O—Au<sub>2</sub>O heterocycle has a chair conformation. The **oxonium trifluoroacetate m 209-210° (dec)** from CHCl<sub>3</sub>/\*C<sub>6</sub>H<sub>6</sub>, and the **oxonium permanganate m 131-131.5° (dec)** from CHCl<sub>3</sub>/hexane (1.2:4.0 v/v) were prepared similarly. [Nesmeyanov et al. *J Organomet Chem* **201** 343 1980, cf Bruce et al. *Inorg Synth* **26** 326 1988.] Oxonium complexes with other phosphine ligands have been similarly prepared and if H<sub>2</sub><sup>17</sup>O is used in the preparations, then it is incorporated into the oxygen of the cation. [Ph<sub>3</sub>PAu(I)]<sub>3</sub><sup>17</sup>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup> has <sup>31</sup>P NMR (36MHz, CD<sub>2</sub>Cl<sub>2</sub>, external 85% H<sub>3</sub>PO<sub>4</sub>) with  $\delta_P$  at 24.0, and <sup>17</sup>O NMR (CD<sub>2</sub>Cl<sub>2</sub> and external H<sub>2</sub>O) with  $\delta_{17O}$  at +19.7 (br s, w/2 = 152Hz). [Yang et al. *Inorg Chem* **32** 1946 1993.]

[Ph<sub>3</sub>PAu(I)]<sub>3</sub>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup> catalyses the Claisen rearrangement of propargyl vinyl ethers, e.g. 1-phenylhept-2-yn-1-yl vinyl ether, to respective aldehydes which are usually reduced *in situ* with NaBH<sub>4</sub> to give the homoallylic alcohol, i.e. 1-phenyl-3-*n*-butyl-3-(2-hydroxyethyl)allene, in high yield with 1mole% of catalyst in CH<sub>2</sub>Cl<sub>2</sub> at room temperature. Chirality is efficiently transferred in the rearrangement. [Sherry & Toste *J Am Chem Soc* **126** 15978 2004]. In the presence of sterically hindered secondary amines, e.g. (iso-Pr)<sub>2</sub>NH or (iso-Pr)CyNH, it catalyses the 5-*exo-dig* cyclisation of formyl alkynes, e.g. 6-formyl-4,4-bis(methoxycarbonyl)hex-1-yne, to 1,1-(bismethoxycarbonyl)-3-formyl-4-exomethylenecyclopentane in CDCl<sub>3</sub>, 70°/3-24 hours, in  $\sim 70\%$  yields [Binder et al. *Org Lett* **10** 1025 2008]. In the presence of O<sub>2</sub> in THF at  $\sim 50^\circ$  [Ph<sub>3</sub>PAu(I)]<sub>3</sub>O<sup>+</sup> [BF<sub>4</sub>]<sup>-</sup> also catalyses the cascade cyclisation/oxidative cleavage of a carbon-carbon triple bond in Z-enynols efficiently [see Ph<sub>3</sub>PAu(I)]<sup>+</sup> [OTf]<sup>-</sup> above, Liu et al. *J Am Chem Soc* **128** 11332 2006].

**(Xantphos)Rh(H)(CO)(PPh<sub>3</sub>)** is prepared by stirring a solution of (PPh<sub>3</sub>)Rh(H)(CO) (100mg, 0.11mmol) and Xantphos (63.6mg, 0.11mmol, see [161265-03-8] in Catalysts—Part 2 below) in \*C<sub>6</sub>H<sub>6</sub> (10ml) at 30° for 4 hours, evaporating the solvent *in vacuo*, washing the residue with MeOH (1ml) and drying *in vacuo* to give the analytically pure complex with the formula C<sub>58</sub>H<sub>48</sub>O<sub>2</sub>P<sub>3</sub>Rh. The complex has IR (CHCl<sub>3</sub>) with  $\nu_{\max}$  at 1996.9vs, 1909.6m cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, C<sub>6</sub>D<sub>6</sub>) has  $\delta$  at 7.82 (apparent q, 4H,  $J = 4.8\text{Hz}$ , ar), 7.66 (m, 6H, ar), 7.53 (apparent q, 4H,  $J = 4.9\text{Hz}$ , ar), 7.11 (dd, 2H,  $J = 7.3, 1.3\text{Hz}$ , CHCHCC), 7.0-6.9 (ar), 6.79 ('d', 4H), 1.48 (s, 3H, CCH<sub>3</sub>), 1.38 (s, 3H, CCH<sub>3</sub>), -9.14 ( $J_{H-P} = 12.2\text{Hz}$ ,  $J_{H-P'} = 18.2\text{Hz}$ ,  $J_{H-Rh} = 1.7\text{Hz}$ ); the <sup>13</sup>C{<sup>1</sup>H}NMR (100MHz, C<sub>6</sub>D<sub>6</sub>) has  $\delta$  at 156.9 (t,  $J = 5.6\text{Hz}$ ), 141.9 (dt,  $J = 4.4, 31.1\text{Hz}$ ), 140.1 (t,  $J = 18.9\text{Hz}$ ), 137.6 (t,  $J = 18\text{Hz}$ ), 135.6 (s), 134.9 (ar), 134.8 (ar), 134.6 (ar), 134.5 (ar), 134.4 (ar), 134.2 (ar), 36.9 (s, C(CH<sub>3</sub>)<sub>2</sub>), 31.0 (s, CH<sub>3</sub>), 24.7 (s,

CH<sub>3</sub>); and the <sup>31</sup>P{<sup>1</sup>H}NMR (121.5MHz, C<sub>6</sub>D<sub>6</sub>, referenced to external 85% H<sub>3</sub>PO<sub>4</sub>) has δ at 42.67 (*J*<sub>P-Rh</sub> = 151.1Hz, *J*<sub>P-P</sub> = 119.1Hz, PPh<sub>3</sub>), 25.65 (*J*<sub>P-Rh</sub> = 127.9Hz, *J*<sub>P-P</sub> = 119.1Hz, Xantphos-P); and the MS has *m/z* at 961 (M-CO), 726 (M-PPh<sub>3</sub>-2H), 698 (M-PPh<sub>3</sub>-CO-2H); HR-MS found *M* is 578.1916 (calc for C<sub>39</sub>H<sub>32</sub>OP<sub>2</sub> is 578.1928). [Kranenburg, van der Burgt, Kramer and van Leewen *Organometallics* **14** 3081 1995.]

**Xantphos-Ru complex** formed *in situ* from equimolar amounts of Xantphos and Ru(PPh<sub>3</sub>)(CO)H<sub>2</sub> in refluxing toluene containing piperidinium acetate is a good catalyst for alkylating active methylene compounds, e.g. *t*-butyl cyanoacetate, and hydroxy compounds, e.g. PhCH<sub>2</sub>OH, to provide α-substituted cyanoacetates, e.g. *t*-butyl 2-cyano-3-phenylpropionate [Slatford, Whittlesey and Williams *Tetrahedron Lett* **47** 6787 2006].

**Xantphos-Pd complexes** formed *in situ* from xantphos and Pd(OAc)<sub>2</sub> or Pd<sub>2</sub>(dba)<sub>3</sub> catalyse the cross-coupling reactions between the amide nitrogen and aryl and heteroaryl halides [Manley & Bilodou *Org Lett* **14** 2433 2004] or *meso*-brominated porphyrins [Gao, Chen and Zhang *Org Lett* **11** 1837 2004] to form the corresponding aryl-N-CO- and *meso*-porphyrin-NCO- in high yields respectively in the presence of a base e.g. CsCO<sub>3</sub>. Similar cross-coupling catalysis with *in situ* Xantphos and Pd(OAc)<sub>2</sub> between thenylbromides and 2-aminopyridines or aminoquinolines has been achieved in high yields [Begouin et al. *Synthesis* 2794 2006]. It should be noted that a detailed study of the mixing of Xantphos and Pd<sub>2</sub>(dba)<sub>3</sub> identified the complexes (**Xantphos**)Pd(**dba**) and **Pd(Xantphos)**<sub>2</sub>. The former is a very active catalyst whereas the latter bis(xanthane based phosphine)<sub>2</sub>Pd(0) is less soluble and inherently less active. It is therefore important, when forming the pre-catalyst, to keep the ratio of xantphos to Pd source less than 1.5:1 for a more effective catalyst involving cross-coupling between a nitrogen atom and an organic halide. [Lingensmith, Strieter, Barder and Buchwald *Organometallics* **25** 82 2006.]

(**Xantphos**)<sub>2</sub> Pd is prepared, in flame dried Schlenk equipment, evacuated/backfilled with argon three times, by stirring xantphos (579mg, 1.00mmol) and Pd<sub>2</sub>(dba)<sub>3</sub> (229mg, 0.25mmol) in toluene (300ml) under argon for 4 hours. This is filtered *via* a cannula into a separate dry flask under argon, concentrated somewhat overnight and any palladium black which may have settled is filtered off and the filtrate is evaporated to dryness. The yellow residue is stirred overnight with toluene (100ml) to remove excess (or unused) dibenzylidene acetone and unreacted Xantphos. The remaining yellow Pd(Xantphos)<sub>2</sub> has **m 164° (dec.)**, is sparingly soluble in most common organic solvents, and is characterised by the correct elemental analyses for C<sub>78</sub>H<sub>64</sub>O<sub>2</sub>P<sub>4</sub>Pd. Its IR (KBr) has *v*<sub>max</sub> at 2924, 2854, 1461, 1398, 1377, 1222 cm<sup>-1</sup>; and MALDI-MS: calcd for C<sub>78</sub>H<sub>64</sub>O<sub>2</sub>P<sub>4</sub>Pd: *theoretical*: 1260.2894 (22.9%), 1262.2909 (63.4%), 1262.2911 (100.0%), 1263.2907 (64.6%), 1264.2914 (77.2%), 1265.2933 (47.5%), 1266.2931 (35.5%), 1267.2949 (25.1%); *Found*: 1260.3405 (24.0%), 1261.3285 (67.4%), 1262.3166 (100.0%), 1263.3162 (73.2%), 1264.3300 (79.3%), 1265.3424 (47.5%), 1266.3491 (35.5%), 1267.3104 (25.1%). [Lingensmith, Strieter, Barder and Buchwald *Organometallics* **25** 82 2006.]

**Zirconocene chloride hydride (bis[cyclopentadienyl]zirconium(IV) hydride chloride, Cp<sub>2</sub>ZrClH) (Schwartz' reagent) [37342-97-5] M 257.9.** It is moisture and light sensitive. Determine its purity by reaction with a slight excess of Me<sub>2</sub>CO whereby the active H reacts to produce Cp<sub>2</sub>ZrClOPr<sup>i</sup> and the integrals of the residual Me<sub>2</sub>CO in the <sup>1</sup>H NMR will show its purity. The presence of Cp<sub>2</sub>ZrH<sub>2</sub> can be determined because it forms Cp<sub>2</sub>Zr(OPr<sup>i</sup>)<sub>2</sub>. For a very active compound, it is best to prepare it freshly from the *dichloride*, see below by reduction with Vitride [LiAl(OCH<sub>2</sub>CH<sub>2</sub>OH)<sub>2</sub>H<sub>2</sub>], the white precipitate is filtered off, washed with tetrahydrofuran, then Et<sub>2</sub>O and dried in a vacuum. Store it dry in the dark. [Carr & Schwartz *J Am Chem Soc* **101** 3521 1979, Negishi & Takahashi *Aldrichimica Acta* **18** 31 1985, Buchwald et al. *Tetrahedron Lett* **28** 3895 1987, Negishi & Takahashi *Synthesis* 1 1988, *Beilstein* **16** IV 1770.] It has been used for functionalising olefins and alkyne [Sun et al. *Org Synth* **71** 83 1992, Negishi & Takahashi *Aldrichimica Acta* **18** 31 1985, Ganem & Franke *J Org Chem* **72** 3981 2007]. It has also been used for mild and selective hydrozirconation of amides to aldehydes [Spletstoser et al. *J Am Chem Soc* **129** 3408 2007].

**Zirconocene dichloride (bis[cyclopentadienyl]zirconium dichloride, Cp<sub>2</sub>ZrCl<sub>2</sub>) [1291-32-3] M 292.3, m 242-245°, 248°.** Recrystallise the dichloride from CHCl<sub>3</sub> or xylene and dry it in a vacuum. <sup>1</sup>H NMR (CDCl<sub>3</sub>) has δ at 6.52 from Me<sub>4</sub>Si. Store it dry in the dark under N<sub>2</sub>. [Reid et al. *Aust J Chem* **18** 173 1965, *Beilstein* **16** IV 1770.]. Together with Zn and CH<sub>2</sub>Cl<sub>2</sub> it is used for methylenation of carbonyl compounds [Tour et al.

*Tetrahedron Lett* **30** 3927 1989], and has been useful for the synthesis of a wide range of early transition-metal complexes and organometallic compounds [Negishi & Takahashi *Aldrichimica Acta* **18** 31 1985, see also selected volumes of Fieser & Fieser's *Reagents for Organic Synthesis*].

**NANO METAL CATALYSTS:** see section on Nano Metals in Chapter 8, "Nanomaterials and Nanotechnology".

## ORGANOCATALYSTS

Organocatalysis has already been discussed in general terms (see Chapter 3), and this section deals with some individual examples that have proved useful.

Organocatalysts are small metal-free organic molecules, or metal-containing compounds but where the metal is not involved in the catalytic process, which efficiently and selectively catalyse a large variety of chemical reactions when used in small non-stoichiometric amounts. Several are commercially available and the purification of some are described here; the preparations of some also are mentioned as these assist in devising methods of purification. The purification of a number of organocatalysts is also described in other chapters. Examples include (-)-*quinine* [see 130-95-0] and (+)-*quinidine* [see 56-54-2] used to catalyse  $\alpha$ -halogenation of carbonyl compounds, intramolecular Michael additions, cyclopropanation of enones, enolates etc,  $\beta$ -lactam synthesis from imines and ketenes,  $\beta$ -lactone synthesis from aldehydes and ketenes, Morita-Baylis-Hillman reactions, hydrophosphonylation of aldehydes, Diels-Alder reactions and desymmetrisation of *meso*-anhydrides; (-)-*cinchonidine* [see 405-71-2] and (+)-*cinchonine* [see 118-10-5] which catalyse additions to prochiral ketenes, desymmetrisation of *meso*-diols and of *meso*-epoxides; and their respective *N*-benzyl salts [see 69257-04-1, and 69221-14-3] are efficient chiral *phase transfer catalysts* (see below); *L*-*proline* [see 147-85-3] which catalyses intramolecular Michael addition, Mannich reactions, inter- and intra- molecular aldol reactions, addition to N=O ( $\alpha$ -aminoxylation/hydroxylation of C=O compounds), addition to nitrones and to N=N double bonds ( $\alpha$ -amination of C=O compounds); *L*-*phenylalanine* [see 63-91-2] which catalyses intramolecular aldol reactions; and the *metal-free Jacobsen's ligand N,N'*-Bis(3,5-di-*tert*-butylsalicylidene)-1,2-*trans*-cyclohexanediamine [see *R,R'*-135616-40-9 or *S,S'*-135616-36-3] which catalyses hydrocyanation (with TMS-CN) of aldehydes. It should be noted that the *Phase Transfer Catalysts* described in a later section of this Chapter are *organocatalysts*. [For further examples see: A. Berkessel and H. Gröger *Asymmetric Organocatalysis* Wiley-VCH, Weinheim 2005, ISBN-13 978-3-527-30517-9, ISBN-10 3-527-30517-3; B. List (Ed) *Topics in current chemistry Vol 291: Asymmetric Organocatalysis* Springer Verlag 2010, ISBN 9783642028144, 9783642028151; which include sections on *Noncovalent organocatalysis based on hydrogen bonding* (Etzenbach-Effers & Berkessel), *Enamine catalysis* (Pihko et al.), *Carbene catalysis* (Moore & Rovis) *Bronsted base catalysts* (Ting et al.), *Chiral ketone and iminium catalysts for olefin epoxidation* (Wong & Shi), *Amine, alcohol and phosphine catalysts for acyl transfer reactions* (Spivey & Arseniyadis), *Secondary and primary amine catalysts for iminium catalysis* (Brazier & Tomkinson), *Lewis acid organocatalysts* (Sereda et al.) and *Chiral Bronsted acids for asymmetric organocatalysis* (Kampen et al.); P.I. Dalko (Ed), *Enantioselective Organocatalysis: reactions and experimental procedures*, Wiley-VCH, Weinheim, 2007, ISBN 978-3-527-31522-2, 3-527-31522-5; M.C. Willis, review on *Enantioselective desymmetrisation*, *J Chem Soc, Perkin Trans 1* 1765 1999.] A whole issue of *Chem Rev (ACS)* [107 513-5841 2007] devoted to *Organocatalysis* with B. List as introductory guest writer has appeared.

***RS*( $\pm$ )-1,1'-Bi(2-naphthol) ( $\pm$  1,1'-binaphthyl-2,2'-diol) [602-09-5] M 286.3, m 214-217 $^{\circ}$ , 216-218 $^{\circ}$ , 217-218 $^{\circ}$ , 218 $^{\circ}$ , pK<sub>Est(1)</sub><sup>20</sup> ~7.5, pK<sub>Est(2)</sub><sup>20</sup> ~11.8; *R*(+)-1,1'-bi(2-naphthol) (*R*(+)-BINOL) [18531-94-7] m 208-210 $^{\circ}$ , 207.5-208.5 $^{\circ}$ , [ $\alpha$ ]<sub>D</sub><sup>20</sup> 0 (c 1, THF), [ $\alpha$ ]<sub>546</sub><sup>25</sup> +50.9 $^{\circ}$  (c 1, THF); *S*(-)-1,1'-bi(2-naphthol) (*S*(-)-BINOL) [18531-99-2] m 207-208 $^{\circ}$ , 208-210 $^{\circ}$ , [ $\alpha$ ]<sub>D</sub><sup>22</sup> -34 $^{\circ}$  (c 1, THF), [ $\alpha$ ]<sub>546</sub><sup>25</sup> +51.3 $^{\circ}$  (c 1, THF). 1,1'-Bi(2-naphthol) is easily prepared by slowly adding a solution of FeCl<sub>3</sub>·6H<sub>2</sub>O (28g, 104mmol, see [10025-77-1]) in H<sub>2</sub>O (60ml) to a vigorously stirred and refluxing suspension of 2-naphthol (14.4g, 100mmol, see [135-19-3]) in H<sub>2</sub>O (600ml) and boiled until the oily drops of 2-naphthol disappear and the *binaphthol* separates out in flakes. Boiling is continued for 10 minutes, and the solid is filtered off, washed with boiling H<sub>2</sub>O and dried first in air then *in vacuo*, to give ( $\pm$ )-BINOL in almost quantitative yield. Recrystallisation from hot EtOH, toluene or \*C<sub>6</sub>H<sub>6</sub> (~150ml, using charcoal) gives colourless glistening plates m 218 $^{\circ}$ . The ( $\pm$ )-*dimethyl ether* [75640-87-8, 75685-**

01-7] has **m** 190° (from AcOH). [A.I.Vogel *Practical Organic Chemistry* Longmans, Green & Co, London, p 639 1948, Pummerer et al. *Ber* **59** 2159 1926, Walder *Ber* **15** 2166 1882.] The FT-IR (Nujol) has  $\nu_{\max}$  at 3486.6, 1617.9, 1508.8, 1322.7, 1252.4, 1177.2, 1126.4, 827.3 and 751.3  $\text{cm}^{-1}$  (note below how it is different from its enantiomers). [*Beilstein* **6 H** 1051, **6 I** 519, **6 II** 1026, **6 III** 5877, **6 IV** 7020.]

**R-** and **S-BINOL** have been obtained by a variety of ways which attest to their usefulness. They have been prepared by resolution of the ( $\pm$ )-BNP-acid (see below), the ( $\pm$ )-BINOL-boron ester (*via* the quinine salt), the ( $\pm$ )-BINOL dibenzoate (*via* pancreas acetone powder), or by oxidative coupling of 2-naphthol in the presence of chiral amines [see Zimmer & Suhrbier *J Prakt Chem* **339** 758 1997]. The preparations involving reduction of the ( $\pm$ )-BNP acid directly with  $\text{LiAlH}_4$ , or *via* the methyl phosphate derivative with  $\text{RedAl}$  [sodium bis(2-methoxyethoxy)aluminium hydride] in 3.4M toluene solution [Truesdale *Org Synth Coll Vol VIII* 46 1993] appear to be most convenient, and do not involve racemisation. The former is described here as it does not require an esterification step. Pure **R-(+)-BNP** acid (115.4g, 331mmol) in dry THF at 0° under  $\text{N}_2$ , is treated with  $\text{LiAlH}_4$  (31.4g, 830mmol) in small portions (exert extreme care here because of the large quantities used — fire hazard) during 1 hour with stirring. A further volume of dry THF (400ml) is added and the mixture is stirred at 25° under  $\text{N}_2$ , for 17 hours, cooled to 0° under  $\text{N}_2$ , and cold 6N aqueous HCl (250ml) is added very carefully. The upper phase is then decanted off, and the lower phase is mixed with 6N aqueous HCl (150ml) and THF (150ml), stirred, and is allowed to settle. The phases are separated and the lower phase is extracted with  $\text{Et}_2\text{O}$ . The combined organic phases are washed with brine, decolorised with charcoal and evaporated *in vacuo*. The residual oil is dissolved in  $^*\text{C}_6\text{H}_6$  (1L, toluene may be used) and evaporated until crystallisation sets in. The solid is collected (90.6g, 96%, **m** 202-207°), which upon crystallisation from  $^*\text{C}_6\text{H}_6$  or toluene gives optically pure BINOL (84.5g, 89%, **m** 207-208°) with (+) or (–)  $[\alpha]_{589}^{25}$  34.3°,  $[\alpha]_{578}^{25}$  37.8°,  $[\alpha]_{546}^{25}$  51.3°,  $[\alpha]_{436}^{25}$  228° (**c** 1.1, **THF**). The FT-IR (Nujol) of the BINOL enantiomers are identical, though different from that of the racemate (see above), and have  $\nu_{\max}$  at 3509.9, 1617.9, 1511.5, 1319.1, 1220.9, 1148.9, 815.6, 749.6 and 566.0  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (1:1  $\text{CDCl}_3/\text{DMSO}-d_6$ , TMS) has  $\delta$  at 7.04 (d, 2H,  $J = 8.8\text{Hz}$ ), 7.20-7.35 (m, 4H), 7.40 (d, 2H,  $J = 8.8\text{Hz}$ ), 7.92 (d, 2H,  $J = 8.8\text{Hz}$ ) and 9.21 (s, 2H). **Racemisation studies** showed that: BINOL (0.1g) is optically stable in an  $\text{H}_2\text{O}$  (10ml)/dioxane (12ml) mixture at 100° under  $\text{N}_2$  during 26 hours; loses 44% and 66% of optical activity in BuOH (10ml) containing KOH (0.71mmol) per BINOL (0.33mmol) at 118° under  $\text{N}_2$  in 13 and 23 hours respectively; and loses 37% and 72% of optical activity of BINOL (100mg) in a 20% aqueous HCl (10ml)/dioxane (12ml) mixture at 100° under  $\text{N}_2$  during 7 and 26 hours respectively. [Kyba et al. *J Org Chem* **42** 4173 1977.]

BINOL and 3,3'-diaryl derivatives act as organocatalysts for Mirata-Baylis-Hillman reactions involving C-H activation, e.g. condensation of a variety of aldehyde with cyclohex-2-en-1-one to provide chiral 2-(1-substituted-1-hydroxymethyl)cyclohex-2-en-1-ones in 40 to 80% yields and “ee” values of ~90 in the presence of  $\text{Et}_3\text{P}$  [McDougal & Schaus *J Am Chem Soc* **125** 12094 2003], they catalyse hydrocyanation (with TMS-CN) of aldehydes with high “ee” at -78° [Holmes & Kagan *Tetrahedron Lett* **41** 7453 and 7457 2000], and the titanium complexes catalyse the stereoselective reduction [with  $(\text{MeO})_3\text{SiH}$ ] of carbonyl compounds [Schiffers & Kaga *Synlett* 1175 1997], as well as C-C forming reactions [Zimmer & Suhrbier *J Prakt Chem* **339** 758 1997] and Diels-Alder reactions [Mikami et al. *Tetrahedron: Asymmetry* **2** 643 1991]. The rare earth La-S-BINOL complex realises an asymmetric nitroaldol reaction with nitromethane in high product yield and enantiomeric enrichment [Sasai et al. *J Am Chem Soc* **115** 10372 1993].

**RS-( $\pm$ )-1,1'-Binaphthyl-2,2'-diylhydrogen phosphate (1,1'-binaphthyl-2,2'-phosphoric acid, BNP)** [36193-63-6] **M** 348.3, **m** 217°,  $\text{pK}^{20}$  0.74; **R-(–)-1,1'-binaphthyl-2,2'-diylhydrogen phosphate** [39648-67-4] **m** 217°,  $[\alpha]_{\text{D}}^{20}$  –608° (**c** 1, MeOH),  $[\alpha]_{\text{D}}^{22}$  –562.7° (**c** 0.97,  $\text{CHCl}_3$ ),  $[\alpha]_{546}^{22}$  –113° (**c** 0.95,  $\text{CHCl}_3$ ); **S-(+)-1,1'-binaphthyl-2,2'-diylhydrogen phosphate** [35193-64-7] **m** 217°,  $[\alpha]_{\text{D}}^{20}$  +608° (**c** 1, MeOH). Recrystallise it from EtOH. Reflux for 3 hours in N NaOH is required to hydrolyse the cyclic phosphate. The **R-enantiomer** has been prepared by a general procedure in which **R-(+)-1,1'-binaphthyl-2,2'-ol** (1.90mmol, see **R-BINOL** [18531-94-7]) in pyridine (7.8ml) is treated with  $\text{POCl}_3$  (250 $\mu\text{l}$ , 2.68mmol) and stirred at ~25° for 3 hours. The mixture is then quenched with  $\text{H}_2\text{O}$  (154 $\mu\text{l}$ ) at 0°, stirred for 1 hour at ~25°, evaporated *in vacuo*, the residue is acidified with 6N HCl (20ml) at 0° then refluxed for 2 hours. After cooling to 0°, the *phosphoric acid* separates, is filtered off, washed with  $\text{H}_2\text{O}$  and dried *in vacuo*. Recrystallisation by dissolving in  $\text{CH}_2\text{Cl}_2$  and pouring it into hexane provides pure crystals of the **R-(–)-acid** (1.58mmol, 83%). [Yamanaka et al. *J Am Chem Soc* **129** 6756 2007, for general procedure see Jacques & Fourquet *Org Synth* **67** 1 1989.] It has IR ( $\text{CHCl}_3$ ) with  $\nu_{\max}$  at 3024, 3015, 1231, 1215, 1203, 1020, 951, 797, 779, 773, 762, 748, 737, 723, 716, 700 and 673  $\text{cm}^{-1}$ ; its  $^1\text{H}$  NMR

(400MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 8.01 (d, 2H,  $J = 8.8$  Hz), 7.94 (d, 2H,  $J = 8.2$  Hz), 7.55 (d, 2H,  $J = 8.8$  Hz), 7.49-7.44 (m, 2H), 7.38 (d, 2H,  $J = 8.2$  Hz), 7.32-7.27 (m, 2H) and 3.22 (s, 1H); and its <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>, CDCl<sub>3</sub> as internal standard with  $\delta$  at 77.0) has  $\delta$  at 147.2, 147.1, 132.2, 131.6, 131.0, 128.3, 127.0, 126.5, 125.5, 121.4 and 120.5. [Jacques et al. *Tetrahedron Lett* 4617 1971, Arnold et al. *Tetrahedron* **24**, 343 1983, Beilstein 6 II 1027.] **Alternatively**, (and more useful as it can be carried out on a 100g scale), the *R*-(-) and *S*-(+)-BNP acids are obtained by optical resolution of the diastereoisomeric (+)-cinchonine salts of the (±)-BNP acid via recrystallisation from hot MeOH. The crude (+)-acid (+)-base salt [which contains 91% of (+)-acid (+)-base and 9% of (-)-acid (+)-base] has  $[\alpha]_{546}^{25} +424^{\circ}$  (c 0.99, MeOH) and the more soluble crude (-)-acid (+)-base salt obtained by evaporation of the mother liquors [which contains 81% of (-)-acid (+)-base and 19% of (+)-acid (+)-base] has  $[\alpha]_{546}^{25} -113^{\circ}$  (c 0.95, MeOH). Pure (+)-acid (+)-base salt  $[\alpha]_{546}^{25} +492^{\circ}$  (c 1, MeOH) and pure (-)-acid (+)-base salt,  $[\alpha]_{546}^{25} -256^{\circ}$  (c 1, MeOH) as obtained from the pure enantiomeric BNP acids and (+)-cinchonine are recrystallised from MeOH/EtOAc and MeOH/Me<sub>2</sub>CO/EtOAc respectively. The BNP acids are obtained from the respective salts (e.g. 77g) in boiling EtOH (e.g. 500ml) and adding 6N aqueous HCl (e.g. 570ml) with vigorous stirring during 30 minutes while keeping the temperature between 75° and 80°. After cooling the BNP acid is collected, washed with H<sub>2</sub>O (e.g. 5 x 90ml), and air dried to give (~50 yield of ~99% ee) almost pure acid (by NMR and HPLC). BNP acids from the filtrates can be purified by recrystallisation from hot MeOH or EtOH and are sparingly soluble in H<sub>2</sub>O. The solubilities at 25° in g/100ml are: 5.7±0.2 enantiomer (EtOH) and 10.3±0.5 racemate (EtOH); and 2.1±0.1 enantiomer (MeOH) and 2.5±0.1 racemate (MeOH). Two recrystallisations from EtOH of the BNP enantiomeric acids obtained did not alter the rotations at 25° which are (+) or (-): 595° ±7 (589nm), 624° ±7 (578nm), 720° ±8 (546nm), 1328 ±15 (436nm), 2050° ±25 (365nm) in MeOH; 574° ±16 (589nm), 602° ±17 (578nm), 694° ±20 (546nm), 1267 ±25 (436nm), 1828° ±40 (365nm) in EtOH. [Jacques & Fouquet *Org Synth Coll Vol VIII* 51 1993, see also Kyba et al. *J Org Chem* **42** 4172 1977].

The BNP acids, and their 3,3'-aryl derivatives, are useful asymmetric organocatalysts in being efficient *Bronsted acid catalysts* for enantioselective Mannich reactions. [Yamanaka et al. *J Am Chem Soc* **129** 6756 2007.]

**2,2'-Bis(diphenylphosphino)-1,1'-binaphthyl (BINAP)** [*RS* see 98327-87-8] **M 622.7, m 283-286°**. [*R*-(+)-see 76189-55-4] **m 241-242°**, [*S*-(-)-see 76189-56-5] **m 241-242°**. In addition to being ligands for Pd, Rh and Ru catalysts, *S*-BINAP also act as a organocatalysts for Mirata-Baylis-Hillman reactions involving C-H activation, e.g. in the condensation of 5-formylpyrimidine and derivatives with methyl acrylate to form methyl β-hydroxy-β-(pyrimidin-5-yl)-α-methylenepropionates with moderate enantiomeric excess at the chiral hydroxy centre [Hayase et al. *Chem Commun* 1271 1998].

### Bode-Rovis-Enders NHC precatalysts

These are a group of 3,4-fused-2-aryl substituted 1,2,4-triazolium salt pre-catalysts, which when converted into the corresponding NHCs (N-heterocyclic carbenes) where C-5 of the heterocyclic ring becomes a carbene carbon atom, are **organocatalysts** that promote a wide variety of organic reactions. The carbene is usually generated *in situ* on addition of diisopropylamine or BTU. Among the reactions that these NHCs catalyse are cyclocondensation reactions, enantioselective azadiene Diels-Alder reactions, aza-Claisen reactions, annulation reactions, formation of trisubstituted indanes *via* catalysed cascade annulations, desymmetrisation of cyclohexadienones, benzoin condensations, Stetter reaction, a<sup>3</sup> to d<sup>3</sup> Umpolung, transesterification, 1,2-additions, etc. [Bode *Speciality Chemicals Magazine* **27(6)** 28 pages 2007, Moor & Rovis *Top Curr Chem* **290** 77-144 2009, Enders, Niemeier and Henseler *Chem Rev* **107** 5606-5655 2007.] Described below are such four chiral organoprecatalysts whose carbenes catalyse highly stereospecific reactions in good to very good yields and with very high enantiomeric excess(e,e), and two related non-chiral salts which also catalyse non-chiral reactions in high yields.

**(4a*R*,9a*S*)-4,4a,9,9a-1-oxa-4-azafluorenone**, **(4a*R*,9a*S*)-indeno[2,1-*b*]-1,4-oxazin-3(2*H*)-one**, [862095-79-2] **M 189.2, m 180-183°(dec)**,  $[\alpha]_{\text{D}}^{23} -17.3^{\circ}$  (c 1.08, MeOH). This is the key chiral intermediate *fluorene-3-one* that uses (+)-*cis*-(1*R*,2*S*)-1-aminoindan-2-ol (see [136030-00-7]) for the preparation of chiral *indeno-triazolo-oxazinium* pre-catalysts. Details of its synthesis are not given here because they have been reported in great detail by Rovis and coworkers [Vora et al. *Org Synth* **87** 350 2010], but the physical and spectroscopic properties, and final purification procedures are provided in order to assess the purity of the commercial product and to purify it further if necessary. Thus the aminoindanol is converted to the -ONa salt with NaH/THF which is reacted with ClCH<sub>2</sub>CO<sub>2</sub>Et to form the *fluorenone*. The final purification is carried out by stirring vigorously the crude light brown solid with hexanes under a reflux condenser in an oil bath at 70° for

2 hours, cooling to  $\sim 25^\circ$ , filtering the solid off and drying it at  $70^\circ$  *in vacuo* (2mm Hg) for 1 hour to give the *fluorenone* as a white solid in 80-86% yield of almost analytical (C, H and N) purity. It has IR with  $\nu_{\max}$  at 3179, 3050, 2910, 1681, 1646, 1484, 1417  $\text{cm}^{-1}$ ;  $^1\text{H}$  NMR (400MHz,  $\text{Me}_2\text{CO}-d_6$ ) with  $\delta$  at 2.97 (d,  $J = 16.8\text{Hz}$ , 1H), 3.23 (dd,  $J = 16.6, 4.9\text{Hz}$ , 1H), 3.89 (d,  $J = 16.2\text{Hz}$ , 1H), 4.05 (d,  $J = 16.2\text{Hz}$ , 1H), 4.57 (t,  $J = 4.6\text{Hz}$ , 1H), 4.82 (t,  $J = 4.0\text{Hz}$ , 1H), 7.23-7.29 (m, 3H), 7.46-7.49 (m, 1H), 8.21 (br s, 1H);  $^{13}\text{C}$  NMR (100MHz,  $\text{Me}_2\text{CO}-d_6$ ) with  $\delta$  at 38.2, 59.4, 67.1, 77.1, 124.7, 125.7, 127.6, 128.4, 140.8, 143.1, 168.5; and HRMS (APCI<sup>+</sup>) has  $m/z$  190.0862, and the calculated value for  $\text{C}_{11}\text{H}_{12}\text{NO}_2$  ( $\text{M}^+$ ) is 190.0863. The enantiomeric (**4aS,9aR**)-**4,4a,9,9a-1-oxa-4-azafluorenone** can be similarly prepared by starting from (-)-*cis*-(1*S*,2*R*)-1-*aminoindan-2-ol* (see [126456-43-7]) and differs only in the sign of the specific rotation.

(**5aR,10bS**)-**5a,10b-Dihydro-2-(2,4,6-trimethylphenyl)-4H,6H-indeno[2,1-b]-1,2,4-triazolo[4,3-d]-1,4-oxazinium chloride monohydrate** [903571-02-8 for anhydrous] **M 385.9, m 217-219 $^\circ$ , 226-230 $^\circ$**  (anhydrous ?),  $[\alpha]_{\text{D}}^{20} -133.5^\circ$  (c 1.00, EtOH),  $[\alpha]_{\text{D}}^{20} -156.0^\circ$  (c 1.00,  $\text{CHCl}_3$ ). Here also details of its synthesis are not given because they have been reported in great detail by Bode and coworkers [see Struble & Bode *Org Synth* **87** 362 2010], but the physical and spectroscopic properties, and final purification procedures are provided in order to assess the purity of the commercial products and to purify further if necessary. The preceding (**4aR,9aS**)-*fluorenone* is treated with  $\text{Me}_3\text{OBF}_4/\text{CH}_2\text{Cl}_2$  ( $\sim 25^\circ/16$  hours) then  $\text{NaHCO}_3$ , and dried in a high vacuum (12 hours) to give a 92-93% yield (95% purity) of (**4aR,9aS**)-**3-methoxy-2,4a,9,9a-tetrahydroindeno[2,1-b][1,4]oxazine** as a pale brown solid with **m 73-75 $^\circ$** ,  $[\alpha]_{\text{D}}^{20} -15.7^\circ$  (c 1.26, EtOH), and consistent IR, NMR and MS. This *methoxy-oxazine* is then added to a solution containing one equivalent of *2,4,6-trimethylphenylhydrazinium chloride*\* in MeOH and a catalytic amount of 4M HCl/dioxane (obligatory) and heated at  $60^\circ$  under a reflux condenser and  $\text{N}_2$  atmosphere for 48 hours. The solvent is removed *in vacuo* and the crude orange solid residue is suspended in EtOAc (must be absolutely dry for high yields; dry by passing through an  $\text{Al}_2\text{O}_3$  drying column in an argon atmosphere) with vigorous stirring at  $90^\circ$  (under  $\text{N}_2$ ) for 30 minutes, cooled to  $0^\circ$ , and the light yellow solid is filtered off, washed with EtOAc and dried at  $\sim 25^\circ$  ( $<1$  mbar, 12 hours) to provide an 81-82% yield of (**Z**)-**2-mesityl-1-[(4aR,9aS)-4,4a,9,9a-tetrahydroindeno-[2,1b][1,4]-oxazin-3-(2H)ylidene]hydrazinium chloride** as a light yellow powder which has **m 200-201 $^\circ$ (dec)**,  $[\alpha]_{\text{D}}^{20} +66.7^\circ$  (c 1.42, EtOH), and consistent IR, NMR and MS. The precatalyst is then obtained by cyclising the hydrazium chloride with triethylorthoformate (8 equivalents) in chlorobenzene (dried over 4Å molecular sieves) and anhydrous HCl (4M in dioxane, 1 equivalent) by stirring under  $\text{N}_2$  at  $120^\circ$  for 1 hour, and the clear brown solution is evaporated *in vacuo*. The brown residue is treated with toluene, heated at  $120^\circ$  for 5 minutes, cooled in ice/water, the white solid is filtered off and the filtrate is evaporated at  $60^\circ$  *in vacuo* and dried at high vacuum ( $<0.75\text{mm}$ ). The brown residue dissolves in toluene at  $60^\circ$  within 3-5 minutes then a white precipitate is formed rapidly. The cooled solution is filtered, the white solid is washed with a little toluene and dried at high vacuum ( $<1$  mbar) to constant weight ( $>24$  hours) to give a 60-64% yield of the desired (**5aR,10bS**)-*oxazinium chloride* pre-catalyst as an analytically pure (C, H and N) white powder with **m 217-219 $^\circ$** . It has IR with  $\nu_{\max}$  at 3435, 3482, 2904, 2853, 1580, 1466, 1222, 1097, 1083, 847, 749, 729, 662  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (400MHz,  $\text{DMSO}-d_6$ ) has  $\delta$  at 212 (s, 6H), 2.37 (s, 3H), 3.16 (d,  $J = 17.0\text{Hz}$ , 1H), 3.50 (dd,  $J = 16.9, 4.8\text{Hz}$ , 1H), 4.99 (t,  $J = 4.4\text{Hz}$ , 1H), 5.08 (d,  $J = 16.0\text{Hz}$ , 1H), 5.26 (d,  $J = 16.0\text{Hz}$ , 1H), 6.12 (d,  $J = 4.0\text{Hz}$ , 1H), 7.21 (s, 2H), 7.33-7.45 (m, 3H), 7.65 (d,  $J = 7.2\text{Hz}$ , 1H), 11.34 (s, 1H); the  $^{13}\text{C}$  NMR (100MHz,  $\text{DMSO}-d_6$ ) has  $\delta$  at 16.9, 20.6, 37.0, 59.7, 61.1, 76.8, 124.0, 125.3, 127.1, 129.2, 129.3, 131.2, 134.8, 136.1, 140.6, 141.3, 144.7, 150.0; and the HRMS (ESI) has  $m/z$  332.1756 and the calculated value for  $\text{C}_{21}\text{H}_{22}\text{N}_3\text{O}$  ( $\text{M}^+$ ) is 332.1757.

The enantiomeric (**5aS,10bR**)-**5a,10b-Dihydro-2-(2,4,6-trimethylphenyl)-4H,6H-indeno[2,1-b]-1,2,4-triazolo[4,3-d]-1,4-oxazinium chloride monohydrate** **M 385.9, m 212-216 $^\circ$ ,  $[\alpha]_{\text{D}}^{20} +158.0^\circ$**  (c 1.00,  $\text{CHCl}_3$ ) can be prepared similarly by starting from the enantiomeric (**4aS,9aR**)-**4,4a,9,9a-1-oxa-4-azafluorenone** and is purified in the same way.

\*The purity of **2,4,6-mesitylhydrazinium chloride** [76195-82-9] **M 186.7** is critical for obtaining pure mesityl-pre-catalysts. The preparation from 2,4,6-trimethylaniline is described in great detail by Struble and Bode [*Org Synth* **87** 362 2010] who emphasise that it is imperative for diazotisation and reduction to be strictly monitored. It is paramount that  $\text{NaNO}_2$  and  $\text{SnCl}_2$  solutions be added at such a rate that the internal temperature is carefully monitored at  $<5^\circ$ . The final purification is by suspending the salt ( $\sim 7.5\text{g}$ ) in absolute EtOH/Et<sub>2</sub>O (1:1, 15ml), and sonicating (85W) for 5 minutes while maintaining the temperature below  $30^\circ$ . The pale orange solid is filtered off using a fine porosity filter, washed with EtOH/Et<sub>2</sub>O (1:1, 3 x 5ml), and dried under high vacuum ( $\sim 25^\circ$ ,  $<0.75\text{mm}$ ) for 12 hours to give analytically pure (C, H and N) **2,4,6-mesitylhydrazinium chloride** (36-40% yield).

Note that the material may be very fine to filter in a short period, in which case it should be dissolved in dry MeOH and concentrated by evaporation to give a more crystalline product that can be filtered more easily. *Alternatively*, it should be collected by centrifugation, decantation of the supernatant, washed with EtOH/Et<sub>2</sub>O (1:1, 3 x 5ml) also by decantation and sonication in the same way in between washes. It has **m 195-197°(dec)** (**167-173°** also reported); it has IR with  $\nu_{\max}$  at 3296, 3002, 2964, 2911, 2691, 1564, 1515, 1479, 846, 830, 575 cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, DMSO-*d*<sub>6</sub>) has  $\delta$  at 2.20 (s, 3H), 2.35 (s, 6H), 6.60 (br s, 1H), 6.86 (s, 2H), 9.27 (variable br s, 3H, NH,NH<sub>2</sub>); the <sup>13</sup>C NMR (100MHz, DMSO-*d*<sub>6</sub>) has  $\delta$  at 17.8, 20.4, 129.0, 134.9, 136.1, 137.9; and HRMS (ESI) has *m/z* 151.1228, and the calculated value for C<sub>9</sub>H<sub>15</sub>N<sub>2</sub> (M<sup>+</sup>) is 151.1223. [Struble and Bode *Org Synth* **87** 362 2010]

**(5aR,10bS)-5a,10b-Dihydro-2-(pentafluorophenyl)-4H,6H-indeno[2,1-b]-1,2,4-triazolo[4,3-d]-1,4-oxazinium tetrafluoroborate** [740816-14-2] **M 467.1, m 223-226°, 235°(dec)**, [ $\alpha$ ]<sub>D</sub><sup>23</sup> -130.8° (c 1.28, MeCN). Details of its synthesis are not given here because they have been reported in great detail by Rovis and coworkers [Vora et al. *Org Synth* **87** 350 2010], but the physical and spectroscopic properties, and final purification procedures are provided in order to assess the purity of the commercial product and to purify it further if necessary. The common intermediate **(4aR,9aS)-4,4a,9,9a-1-oxa-4-azafluorenone** (1.0 equivalent see [862095-79-2] above) is added with stirring, and under anhydrous conditions in an argon atmosphere, to a mixture of Me<sub>3</sub>OBF<sub>4</sub> (1.0 equivalent *and no more*, care **Toxic**) in CH<sub>2</sub>Cl<sub>2</sub> (containing <20ppm of H<sub>2</sub>O, by filtering through activated Al<sub>2</sub>O<sub>3</sub> under argon) and stirred at ~25° until the homogeneous reaction is complete (1-2 hours by <sup>1</sup>H NMR). Pentafluorophenylhydrazine (1.0 equivalent) is then added to the mixture under argon, stirred until the reaction is complete (~4 hours, when <sup>1</sup>H NMR signals of activated amidate disappear), the mixture is evaporated *in vacuo*, dried at 100°/2mm; then chlorobenzene (as solvent) followed by triethyl orthoformate (2.5 equivalents) are added and heated under reflux to 130° with stirring (open to the atmosphere) for 24 hours. A second portion of triethyl orthoformate (2.5 equivalents, syringe) is added and the mixture is stirred for a further 24 hours, and this process is repeated once more. Excess of toluene is added, the slurry is filtered, and the solid is washed with toluene and hexanes. It is then triturated with EtOAc and MeOH (4:1), filtered through a medium frit funnel, washed with cold EtOAc to give the **oxazinium tetrafluoroborate** pre-catalyst which, when dried at 100°/2mm to constant weight (~1 hour, 61-64% yield), gave an off white solid that analysed as the ~0.5 hydrate. It has IR with  $\nu_{\max}$  at 3147, 3106, 3028, 2967, 1595, 1530, 1517, 1487, 1461, 1056, 1046, 998 cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, Me<sub>2</sub>CO-*d*<sub>6</sub>) has  $\delta$  at 3.28 (d, *J* = 17.2Hz, 1H), 3.55 (dd, *J* = 17.1, 4.9Hz, 1H), 5.19 (t, *J* = 4.5Hz, 1H), 5.25 (d, *J* = 16.4Hz, 1H), 5.39 (d, *J* = 16.4Hz, 1H), 6.33 (d, *J* = 4.0Hz, 1H), 7.34 (t, *J* = 7.3Hz, 1H), 7.43 (q, *J* = 7.4Hz, 2H), 7.63 (d, *J* = 7.6Hz, 1H), 11.09 (br s, 1H); the <sup>13</sup>C NMR (100MHz, Me<sub>2</sub>CO-*d*<sub>6</sub>) has  $\delta$  at 37.9, 60.8, 63.5, 78.2, 125.2, 126.4, 128.1, 130.4, 136.2, 141.7, 147.1, 152.5; and HRMS (APCI<sup>+</sup>) has *m/z* 380.0816, and the calculated value for C<sub>18</sub>H<sub>11</sub>N<sub>3</sub>OF<sub>5</sub> (M<sup>+</sup>) is 380.0817. [Vora et al. *Org Synth* **87** 350 2010].

The enantiomeric **(5aS,10bR)-5a,10b-Dihydro-2-(pentafluorophenyl)-4H,6H-indeno[2,1-b]-1,2,4-triazolo-[4,3-d]-1,4-oxazinium tetrafluoroborate** can be prepared and purified in a similar way but starting with the enantiomeric **(4aS,9aR)-4,4a,9,9a-1-oxa-4-azafluorenone**.

**2-Pentafluorophenyl-6,7-dihydro-5H-pyrrolo[2,1-c][1,2,4]triazol-2-ium tetrafluoroborate** (**2-pentafluoro-2,5,6,7-tetrahydro[2,1-c][1,2,4]-tetrazol-4-ium tetrafluoroborate**), [862095-91-8] **M 363.0, m 242-245°, 245°**. Details of its synthesis are not given here because they have been reported in great detail by Rovis and coworkers [Vora et al. *Org Synth* **87** 350 2010], but the physical and spectroscopic properties, and final purification procedures are provided in order to assess the purity of the commercial product and to purify it further if necessary. Starting with 2-pyrrolidinone (1.0 equivalent, [616-45-5]) which is converted into 2-methoxy-3,4-dihydro-5H-pyrrolidinium tetrafluoroborate by reaction with trimethyloxonium tetrafluoroborate (1.0 equivalent) in CH<sub>2</sub>Cl<sub>2</sub> under argon with stirring at ~25° until homogeneous, then pentafluorophenylhydrazine (1.0 equivalent, [828-73-9]) is added with stirring for 4 hours (when amidate signals disappear in the <sup>1</sup>H NMR spectrum), the mixture is evaporated *in vacuo*, chlorobenzene and triethyl orthoformate (2.0 equivalents, syringe) are added and heated at 130° for 24 hours under a reflux condenser open to the atmosphere. Addition of triethyl orthoformate (2.0 equivalents, syringe) and heating is repeated for 24 hours, followed by evaporation to dryness *in vacuo*, excess of toluene is added to the residue and the slurry is filtered, and the solid is washed with toluene and hexanes, then with EtOAc. Final purification is by washing well with

cold EtOAc, and the off-white solid is thoroughly dried at 100° *in vacuo* (2mm) to constant weight (~1 hour) to give the *pre-catalyst tetrafluoroborate* in 74-76% yield. It has IR with  $\nu_{\max}$  at 3145, 3097, 2983, 1655, 1604, 1524, 1499, 1028, 994, 875  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (400MHz,  $\text{Me}_2\text{CO}-d_6$ ) has  $\delta$  at 3.00 (ddd,  $J = 15.0, 7.7, 7.7\text{Hz}$ , 2H), 3.42 (t,  $J = 8.0\text{Hz}$ , 2H), 4.76 (t,  $J = 8.0\text{Hz}$ , 2H), 10.19 (s, 1H); the  $^{13}\text{C}$  NMR (100MHz,  $\text{Me}_2\text{CO}-d_6$ ) has  $\delta$  at 22.6, 27.6, 49.5, 137.8, 140.4, 142.9, 143.3, 144.5, 145.6, 145.8, 165.8; and HRMS (APCI+) has  $m/z$  276.0556, and the calculated value for  $\text{C}_{11}\text{H}_7\text{N}_3\text{F}_5$  ( $\text{M}^+$ ) is 276.0555. [Vora et al. *Org Synth* **87** 350 2010].

**2-Mesityl-6,7-dihydro-5H-pyrrolo[2,1-c][1,2,4]triazol-2-ium chloride (2-mesityl-2,5,6,7-tetrahydro[2,1-c][1,2,4]-tetrazol-4-ium chloride), M 263.8, m 202-205°.** This is prepared and purified in much the same way as the preceding salt except for using triethyloxonium chloride for preparing the ethoxy pyrrolidinium chloride followed by reaction with mesitylhydrazinium chloride instead of perfluorophenylhydrazine.

**Boron trifluoride diethyl etherate [109-63-7] M 141.9, m -58°, b 46°/10mm, 126-129°/atm,  $d_4^{25}$  1.15,  $n_D^{20}$  1.344.** If the liquid has darkened due to oxidation by air it can be redistilled to give a colourless liquid. It is recommended that excess dry  $\text{Et}_2\text{O}$  (10ml) be added to the etherate (500ml) and distilled from  $\text{CaH}_2$  (2g) in an all glass still at (b 46°/10mm) which removes volatile acids and avoids bumping [Zwiefel & Brown *Organic Reactions* **13** 28 1963]. It is a fuming liquid which hydrolyses in moist air forming highly toxic fluorides on decomposition. Store it under  $\text{N}_2$  or argon and work in an efficient fume cupboard. It is a versatile *Lewis acid* for catalysis, acetylation, alkylation, dehydration, condensations, Beckmann rearrangements and polymerisation reactions. It catalyses the synthesis of cyclopentyl- and cycloheptyl[*b*]indoles from aryl cyclopropyl ketones via [3+2] cycloaddition [Venkatesh et al. *Eur J Org Chem* 5378 2006], and is a mild and efficient catalyst for Beckmann rearrangement in anhydrous MeCN [An et al. *Chin J Chem* **29** 947 2011]. [Beilstein **1** IV 1321.]

#### CBS (Corey-Bakshi-Shibata) Catalysts:

***R*-(+)-2-Methyl-CBS-oxazaborolidine (R-1-methyl-3,3-diphenyltetrahydro-1H,3H-pyrrolo[1,2-c][1,3,2]-oxazaborole,  $\alpha,\alpha$ -diphenyl-D-prolinolmethylboronic acid cyclamide) [112022-83-0] M 277.2, m 74-87°, 79-81°, 85-95°,  $[\alpha]_D^{22} +76.8^\circ$  (c 1, toluene) and *S*-(-)-2-Methyl-CBS-oxazaborolidine [112022-81-8] M 277.2, m 85-95°.** These CBS (Corey-Bakshi-Shibata) catalysts effect the asymmetric reduction of *pro*-chiral ketones, the stereoselective synthesis of  $\alpha$ -amino acids and  $\alpha$ -hydroxy acids, the preparation of ferrocenyl diols and propargyl alcohols. The catalyst is usually best prepared *in situ* from the respective  $\alpha,\alpha$ -diphenyl- (*R* or *S*)-prolinol [22348-32-9 or 79815-20-6] and a slight excess of trimethylboroxine [823-96-1]. Their suitability is determined by capillary GC [DB-1, 200°, retention time 4.9 minutes] and  $^1\text{H}$  NMR, where possible impurities have signals at  $\delta$  ( $\text{CDCl}_3$ ) for starting prolinol at ~4.3(t), trimethylboroxine at 0.45(s), the B—OH intermediate at 0.35 to -0.50 multiple B-Me singlet's, and/or the ring-B-O-B(Me)-OH intermediate at -0.25 B-Me brs. An analytical sample is obtained by dissolving it in toluene, filtering, then evaporating at 50°/0.001mm to give the catalyst as a white solid. Store it under dry  $\text{N}_2$  or argon. The IR ( $\text{CCl}_4$ ) has  $\nu_{\max}$  at 1000, 1235, 1310, 1330, 1440, 2880, 2960  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR (0.2M in  $\text{CDCl}_3$ ) has  $\delta$  at 7.65-7.15 (m, 10 H, ArH), 4.4 (dd,  $J = 5.8, 10.0\text{Hz}$ , 1 H, C3a-H), 3.45-3.30 (m, 1 H, C6-H), 3.15-3.00 (m, 1 H, C6-H), 1.90-1.55 (m, 3 H, C4-H, C5-H<sub>2</sub>), 0.95-0.75 (m, 1 H, C4-H), 0.40 (s, 3 H, B-CH<sub>3</sub>). They are also available commercially as 1M solutions in toluene or tetrahydrofuran. [Mathre et al. *J Org Chem* **56** 571 1991, Corey et al. *J Am Chem Soc* **124** 3808 2002, Ryu & Corey *J Am Chem Soc* **126** 8106 2004, Ryu et al. *J Am Chem Soc* **124** 9992 2002, Hu et al. *J Am Chem Soc* **126** 13708 2004, Quallich & Woodall *Tetrahedron Lett* **34** 785 1993.]

***R*-(+)- and *S*-(-)-2-*o*-Tolyl-CBS-oxazaborolidine {*R*- and *S*-3,3-diphenyl-1-*o*-tolyltetrahydro-1H,3H-pyrrolo[1,2-c][1,3,2]-oxazaborole} [*R*-865812-10-8; *S* 463941-07-3] M 353.3,  $[\alpha]_D^{20} +18^\circ$  and  $-18^\circ$  (c 1, toluene).** These Corey-Bakshi-Shibata (CBS) oxazaborolidines are available commercially as 0.5M solutions in toluene ( $d_4^{25}$  0.9009g/ml). They produce chiral *Lewis acids* when protonated with trifluoromethanesulfonamide and are used in enantioselective Diels-Alder reactions. They are prepared and purified as for the 2-methyl-CBS-oxazaborolidines above except that they are obtained as liquids on evaporation of the toluene solutions. [For applications see Ryu & Corey *J Am Chem Soc* **125** 6388 2003, Stemmler *Synlett* 997 2007, Corey et al. *J Am Chem Soc* **109** 7925 1987, Cho & Kim *Tetrahedron: Asymmetry* **12** 2043 2001, Cho et al. *J Chem Soc, Perkin Trans I* 1204 2001, Lebsack et al. *J Am Chem Soc* **123** 4851 2001, and references therein.]

**2-Methyl-CBS-oxazaborolidines derived from 1*S*,2*R*-(-) [126456-43-7] and 1*R*,2*S*-(+)-*cis*-1-Amino-2-hydroxyindane [136030-00-7]** are best prepared *in situ* by reaction of *cis*-1-aminoindan-2-ol (1*S*,2*R*

see [126456-43-7]; 1*R*,2*S* see [136030-00-7]) with trimethylboroxine (see [823-96-1]) in a solvent (usually toluene), and they are sensitive to moisture and air. They catalyse the reduction of ketones and  $\alpha$ -haloketones to their respective alcohols by  $\text{BH}_3$  or some of its complexes, e.g.  $\text{BH}_3\text{-DMS}$ ,  $\text{BH}_3\text{-THF}$ , with high stereospecificities (ee above 80%) at room temperature or below because of the steric effects of the methyl and indane groups around the heterocyclic boron atom. In a typical preparation, trimethylboroxine (0.19ml, 1.34mmol) is added to a solution of pure 1*R*,2*S*-1-aminoindan-2-ol (2.01mmol) in dry toluene (10ml), stirred at  $\sim 25^\circ$  for 30 minutes, more toluene (10ml) is added and the solution is concentrated by distillation to *ca* 2ml. The procedure is repeated twice then the toluene is removed *in vacuo* to give the *CBS-catalyst* as a white solid. Dry THF ( $\sim 2$ ml) is added to produce an approximately 1.0M solution of the catalyst in THF for later reactions. The solution should be stored under dry  $\text{N}_2$  or argon and is stable for about 48 hours. For reduction reactions, the borane hydride (1.83mmol) is added to a solution of the chiral 2-methylaminoindanol (0.17ml, 0.166mmol) in the desired solvent (e.g. toluene) (2ml) stirred under  $\text{N}_2$  at  $\sim 25^\circ$  for 30 minutes, cooled to  $0^\circ$  and the ketone (1.67mmol) in the desired solvent (e.g. toluene, 1ml) is added *via* a cannula, stirred at  $\sim 25^\circ$  for 30 minutes, quenched with MeOH (5ml),  $\text{H}_2\text{O}$  (20ml) is added, the MeOH is removed *in vacuo*, and the resulting optically active alcohol is isolated from the aqueous solution. [Gilmore & Jones *Tetrahedron: Asymmetry* **14** 2115 2003, Hong et al. *Tetrahedron Lett* **36** 6631 1994

**4-Dimethylaminopyridine (DMAP)** [1122-58-3] **M 122.2, m 108-109 $^\circ$ , 112-113 $^\circ$ , b 191 $^\circ$ /atm (under  $\text{N}_2$ ),  $\text{pK}^{5.4}$  10.14,  $\text{pK}^{20}$  9.70,  $\text{pK}^{25}$  9.61,  $\text{pK}^{35}$  9.33 (in  $\text{H}_2\text{O}$ ).** Recrystallise DMAP from toluene [Sadownik et al. *J Am Chem Soc* **108** 7789 1986]. DMAP can also be purified by recrystallisation from EtOH or  $^*\text{C}_6\text{H}_6$  and dried *in vacuo*. It is prepared easily from suitable 4-substituted pyridines, e.g. 4-chloropyridine and  $\text{Me}_2\text{NH}$  in EtOH (2 hours at  $100^\circ$ , or in a sealed tube for 12 hours at  $100^\circ$ ,  $115\text{-}130^\circ$ , Wibaut & Brockman *Rec Trav Chim Pays Bas* **80** 309 1961), or in very high yield from 4-trimethylsilyloxy pyridine with  $\text{Me}_2\text{NH}/\text{HgCl}_2$  (48 hours at  $120^\circ$ , Vorbrüggen *Angew Chem Int Ed Engl* **11** 305 1972). This organocatalyst is a strong base and promotes a large variety of acylation reactions, e.g. of *m*-chloroaniline and of benzyl alcohol at relative rates that are  $3.14 \times 10^6$  and  $3.45 \times 10^8$  compared with those of the uncatalysed reactions. It is useful industrially. The acylating agent is an acid chloride or an acid anhydride. Typically, catalytic amounts of DMAP are used together with one equivalent of a tertiary base, e.g.  $\text{Me}_3\text{N}$ , to neutralise the anion formed, i.e.  $\text{Cl}^-$  from the acid chloride or  $\text{AcO}^-$  from the anhydride, and to avoid protonating the catalyst. [Scriven *Chem Soc Rev* **12** 129 1983.] Its *picrate* has **m 108 $^\circ$**  (EtOH), and its *tetrafluoroborate salt* has **m 126 $^\circ$**  (from  $\text{MeCN}/\text{Et}_2\text{O}$ ). The UV has  $\lambda_{\text{max}}$  ( $\epsilon$ ) at 261nm ( $18,300 \text{ cm}^2\text{.mole}^{-1}$ ) for free base at pH  $>12.0$ , and 280.5nm ( $19,600 \text{ cm}^2\text{.mole}^{-1}$ ) for the mono-cation at pH  $<7.0$  in  $\text{H}_2\text{O}$  (Essery & Schofield *J Chem Soc* 3939 1961); and the IR ( $\text{CHCl}_3$ ) has  $\nu_{\text{max}}$  ( $\epsilon_{\lambda}$ ) at 2840 (sh, 45), 1448 (95, Me bend), 1380 (155, CN str), 1346 (35), 1179 (35,  $\text{Me}_2$  rock), 1108 (30,  $\text{Me}_2$  rock), 1063 (35,  $\text{NMe}_2$ , C-N-C stretch) and 951 (55,  $\text{NMe}_2$ , C-N-C stretch)  $\text{cm}^{-1}$  (Katritzky & Jones *J Chem Soc* 3674 1959). [*Beilstein* **22** II 341, **22** III/IV 4101, **22/9** V 112.]

DMAP is an efficient *desymmetrisation* catalyst. [Willis *J Chem Soc, Perkin 1* 1765 1999.] Thus it allows the use of an achiral or *meso* molecule, e.g. 3-substituted glutaric anhydride to react with chiral alcohols e.g. *R*-1-(1'-naphthyl)ethanol to produce (1'*R*,3*R*)-1(1'-naphthyl)ethyl methyl 3-substituted-pentanoate di-esters (after treatment with diazomethane), which are now chiral at C-3. The achiral carbon atom at C-3 in the glutaric anhydride becomes chiral in the pentanoate (mono-ester of glutarate). [Theisen & Heathcock *J Org Chem* **53** 2374 1988, Theisen & Heathcock *J Org Chem* **58** 142 1993.]

§ It is commercially available in 0.04mmol impregnated tablets, and a polystyrene supported version (PS-DMAP) with a loading of  $\sim 3.0$ mmol/g is also available for recyclable acylation purposes.

***N,N'*-Dimethylethylenediamine [DMEDA, bis(methylamino)ethane]** [110-70-3] **M 88.2, b 110-112 $^\circ$ /750mm, 119 $^\circ$ /atm,  $d^{25}$  0.819,  $n_D^{20}$  1.431,  $\text{pK}_1^0$  8.30,  $\text{pK}_2^0$  10.89,  $\text{pK}_1^{25}$  6.80 (7.01),  $\text{pK}_2^{25}$  9.88 (10.03).** It has been prepared in several ways, but most commonly when ethylene dibromide (1mol),  $\text{MeNH}_2$  (5mols) and  $\text{H}_2\text{O}$  (*ca* 3mols) are boiled under reflux for 15 hours. A 32% aqueous solution of NaOH is added, excess  $\text{MeNH}_2$  is distilled out, and the rest of the liquid is distilled off. The diamine in the distillate is salted out with solid NaOH, dried ( $\text{Na}_2\text{CO}_3$ , or KOH), filtered and fractionally distilled (in *ca* 50% yield) until the NMR indicates that it is free from impurities. Store it in sealed containers under  $\text{N}_2$  or argon as it is a strong base. The *N,N'*-bis(*p*-toluenesulfonamide) has **m 164 $^\circ$**  (from AcOH), *N,N'*-bis(benzenesulfonamide) has **m 129-131 $^\circ$**  (from EtOH and AcOH, or EtOAc), the *N,N'*-bis(benzamide) has **m 177-178 $^\circ$**  (from  $^*\text{C}_6\text{H}_6$ ), the *monopicrate* has **m 140 $^\circ$**  (plates

from EtOH or Me<sub>2</sub>CO) and the *dihydrochloride* decomposes at **m 235°**. [Kermack & Wight *J Chem Soc* 1425 1935, Boon *J Chem Soc* 311 1947, Woodburn & O'Gee *J Org Chem* 17 1241 1953, Bauer *J Am Chem Soc* 78 1945 1956.]

In the presence of resublimed *tert*-BuOK (1.5mmol), DMEDA (0.1mmol) *catalyses* cross-coupling between aryl and substituted aryl halides (0.5mmol) with arenes, e.g. \*benzene, naphthalene (0.5mmol) at 80°/~4 hours, to provide the corresponding biaryls in high yields (~70 to >90%). *Ortho* substitution leads to decreased yields, and replacing KOBu<sup>t</sup> with other bases (e.g. KOH, Na<sub>2</sub>CO<sub>3</sub>, KOAc, *tert*-BuONa, or *tert*-BuOK, *tert*-BuOLi) results in no reaction or extremely low yields. Of the amines studied DMEDA is the most effective catalyst and evidence is presented that a radical reaction is involved. It is the first example of a cross-coupling reaction which does not require a metal. [Liu et al. *J Am Chem Soc* 132 16737 2010, *Beilstein* 4 H 250, 4 I 425, 4 II 689, 4 III 512, 4 IV 1171.]

***R*-(+)-Indoline-2-carboxylic acid** [98167-06 -7] **M 163.2, m 177°(dec)**, [α]<sub>D</sub><sup>20</sup> +11.0° (c 0.3, MeOH). Purify it as for the *S*-enantiomer below. The *R*-(+)-*hydrochloride* has [172152-19-1]. It is a proline based organocatalyst which promotes the enantioselective formation of cyclopropanes by reaction between 2-(dimethyl-λ<sup>4</sup>-sulfanylidene)-1-phenyl-ethanone (for stable sulfonium ylides see Ratts & Yao *J Org Chem* 31 1185 1966) and but-2-enals in high yields and very high stereoselectivity involving Direct Electrostatic Activation (DEA) [Kunz & Mac Millan *J Am Chem Soc* 127 3240 2005.]

***S*-(-)-Indoline-2-carboxylic acid** [79815-20-6] **M 163.2, m 177°(dec)**, [α]<sub>D</sub><sup>20</sup> -114° (c 1, N HCl). It is purified as the racemate (see [78348-24-0] in Heterocyclic Compound, Chapter 4). The *S*-(-)-*hydrochloride* [82923-76-0] **M 199.7, m 133°(dec)**, [α]<sub>D</sub><sup>20</sup> -70° (c 1, EtOH) crystallises from Et<sub>2</sub>O/propan-2-ol. It is a proline based organocatalyst which promotes the enantioselective formation of cyclopropanes by reaction between 2-(dimethyl-λ<sup>4</sup>-sulfanylidene)-1-phenyl-ethanone (for stable sulfonium ylides see Ratts & Yao *J Org Chem* 31 1185 1966) and but-2-enals in high yields and very high stereoselectivity involving Direct Electrostatic Activation (DEA) [Kunz & Mac Millan *J Am Chem Soc* 127 3240 2005.]

**α-Methyl-L-proline (*S*-2-methylproline)** [42856-71-3] **M 129.2, m 248-252°(dec)**, [α]<sub>D</sub><sup>25</sup> -75° (c 2.0, MeOH), **and (*R*-2-methylproline)** [58123-62-9]. This derivative catalyses reactions similar to those of proline. The general three-step synthesis from proline devised by Seebach and co-workers provides enantiomerically pure compound [*Helv Chim Acta* 64 2704 1981, *J Am Chem Soc* 105 5390 1983, *Org Synth* 72 62 1993]. It has been prepared on a 20-40g scale starting from *S*- or *R*- proline which is converted into **2*S*,5*S*- (or 2*R*,5*S*)- 2-*tert*-butyl-1-aza-3-oxabicyclo[3.3.0]octan-4-one** (*tert*-BuCHO, CF<sub>3</sub>CO<sub>2</sub>H as catalyst, pentane reflux for 114 hours with removal of H<sub>2</sub>O) in 67-74% yield, which is methylated to **2*S*,5*S*- (or 2*R*,5*S*)- 2-*tert*-butyl-5-methyl-1-aza-3-oxabicyclo[3.3.0]octan-4-one** [86046-11-9] (MeI, *iso*-Pr<sub>2</sub>NLi/THF/-78°) with complete retention of configuration at C-5 in 93-95% yield; and hydrolysed (1.3M HCl/reflux for 1 hour, or 15% aqueous HBr/~25° for 8 hours) to give the respective *2-methyl-proline hydrochloride* which is purified through a Dowex 50W x 8 (H<sup>+</sup> form) column by washing with H<sub>2</sub>O, then eluting with 3N aqueous NH<sub>3</sub> and evaporating to give the free *2-methylproline* in 85-90% yield. The last step is apparently quite tedious to carry out and is time consuming. The following procedure is more convenient and gives better than 90% yields of pure *2-methylproline*. The above butyl-methyl-aza-oxa-bicyclooctanone (13.4mmol) in MeOH/H<sub>2</sub>O (6:1, 35ml) is treated with Silica gel 200-400 mesh, 60Å (3.0g) and stirred overnight. The silica gel is filtered off, washed with MeOH, the combined washings and filtrate are evaporated *in vacuo*, the pale yellow residue is dissolved in CHCl<sub>3</sub>/MeOH (20:1), filtered through a 0.45μ filter disk (Acrodisk<sup>®</sup> CR PTFE), evaporated to dryness and the residue is triturated with Et<sub>2</sub>O and dried *in vacuo*. Although it may be yellow in colour, it is almost spectroscopically pure, and can be purified further by dissolving in MeOH, treating with activated charcoal (5% w/w), filtering the whole suspension through Celite and evaporating. It can be crystallised by slow evaporation of the CHCl<sub>3</sub>/MeOH solution, or from an MeOH/Et<sub>2</sub>O mixture to give a first crop (50-70% yield) of white crystals leaving remaining *2-methylproline* and traces of yellow impurities in the mother liquors. Additional product crystallisation can be achieved by vapour diffusion of the supernatant against Et<sub>2</sub>O in a closed desiccator. It also crystallises nicely from MeOH/EtOAc in colourless plates, unlike many α-amino acids. [Genin et al. *Tetrahedron Lett* 35 4967 1994.]

*S*-2-Methylproline with  $[\alpha]_D -72.1^\circ$  (c 1.0, MeOH) was shown to have an enantiomeric excess “ee” of  $99.0 \pm 0.5\%$  by chiral capillary gas chromatography. For this purpose the amino acid (5mg) is heated with *iso*-propyl isocyanate (0.3ml) in  $\text{CH}_2\text{Cl}_2$  (0.3ml) at  $100^\circ$  for 15 minutes, cooled, excess of volatiles are removed by a strong stream of dry air, the residue is dissolved in  $\text{Et}_2\text{O}$  (1ml), filtered through cotton wool and  $10\mu\text{l}$  applied to the chiral capillary column. A Chirasil-Val fused-silica capillary column of Machery-Nagel (25m, 0.4mm) in a Carlo-Erba-Frautovap 4160 HR GC unit has been used. The column was heated at  $160^\circ$  for 15 minutes then raised by  $2^\circ/\text{min}$  up to  $200^\circ$  (Argon as carrier gas ?). The amino acid has  $^1\text{H}$  NMR (200MHz,  $\text{D}_2\text{O}$ , HDO at d 4.80) with  $\delta$  at 1.52 (s, 3H,  $\text{CH}_3$ ), 1.75-2.40 (m, 4H, 3- $\text{CH}_2$  + 4- $\text{CH}_2$ ), and 3.20-3.45 (m, 2H, 5- $\text{CH}_2$ ); and its  $^{13}\text{C}\{^1\text{H}$  decoupled; NMR (50MHz,  $\text{D}_2\text{O}$ ) has  $\delta$  at 179.87 ( $\text{CO}_2\text{H}$ ), 73.2 (C-2), 48.06 ( $\text{CH}_3$ -2), 38.27 (N-C-5), 25-85 (C-3) and 23.99 (C-4). [Beck et al. *Org Synth* **72** 62 1993].

*S*-2-Methylproline hydrobromide [42856-71-3], prepared by hydrolysis with 15% aqueous HBr (see above) evaporating and triturating with dry  $\text{Et}_2\text{O}$  as colourless crystals (80%), has **m** 136-137<sup>o</sup>(dec),  $[\alpha]_D^{20} -28.9^\circ$  (c 1.0, MeOH) and its  $^1\text{H}$  NMR (200MHz,  $\text{D}_2\text{O}$ , HDO at d 4.70) has  $\delta$  at 1.60 (s, 3H,  $\text{CH}_3$ ), 1.80-2.45 (m, 4H, 3- $\text{CH}_2$  + 4- $\text{CH}_2$ ), and 3.36 (t, 2H, J = 7 Hz, 5- $\text{CH}_2$ ) [Seebach et al. *J Am Chem Soc* **105** 5390 1983, Overberger & Jon *J Polym Sci, Polym Sci Ed* **15** 1413 1977].

$\alpha$ -Methyl-DL-proline (*RS*-2-methylproline) [68078-09-1] **M** 129.2, **m** 263-264.5<sup>o</sup>. The racemic 2-methylproline was prepared by hydrolysing 5-(3-chloropropyl)-5-methylhydantoin (2g, 10mmol) with  $\text{Ba}(\text{OH})_2$  (6.3g, 20mmol) and  $\text{H}_2\text{O}$  (50ml) in a pressure bomb at  $160^\circ$  for 30 minutes, cooling, adjusting the pH to 6 with 6N  $\text{H}_2\text{SO}_4$ , filtering off the  $\text{BaSO}_4$ , applying the filtrate through a column of Amberlite IRC 120 ( $\text{H}^+$  form), washing with  $\text{H}_2\text{O}$ , eluting with 4N  $\text{NH}_4\text{OH}$ , and evaporating the eluate to give a white solid m 252-258<sup>o</sup> (1.2g, 90%). Recrystallisation of the solid from MeOH/ $\text{Et}_2\text{O}$  gave analytically pure *RS*-2-methylproline m 263-264.5<sup>o</sup>. It has IR (KBr) with  $\nu_{\text{max}}$  at 3450 and 3200 (OH and NH), and 1600 (C=O)  $\text{cm}^{-1}$ ; and its  $^1\text{H}$  NMR (60MHz,  $\text{CD}_3\text{OD}$ , TMS) has  $\delta$  at 1.6 (s, 3H,  $\text{CH}_3$ ), 1.9 (m, 4H, 3- $\text{CH}_2$  + 4- $\text{CH}_2$ ), and 3.3 (m, 2H, 5- $\text{CH}_2$ ). [Ellington & Honigberg *J Org Chem* **39** 104 1974]. It should be possible to prepare it by Seebach's procedure (see preceding entry) and starting with *RS*-proline.

Tetra-*n*-butylammonium fluoride (TBAF) [anhydrous, 1 M in THF 429-41-4; 3  $\text{H}_2\text{O}$  87749-50-6; hydrate 22206-57-1] **M** 261.5 (anhydrous), 315.5 (3  $\text{H}_2\text{O}$ ), **m** liquid (anhydrous), 62-63<sup>o</sup> (3  $\text{H}_2\text{O}$ ), 37<sup>o</sup> (18  $\text{H}_2\text{O}$ ). It is a powerful catalyst for the regioselective ring opening of epoxides with thiols [Albanese et al, *Synthesis* 34 1994], the addition reactions of trimethylallylsilane [Majetich et al. *J Org Chem* **51** 1745 1986], Michael, Knoevenagel, aldol condensations, and by acting as a base in organic synthesis [Clark *Chem Rev* **80** 429 1980]. It is a mild and highly efficient source of nucleophilic fluoride ions, causing substitution of halogen or tosyl groups by fluorine. TBAF is also used for cleaving silyl ethers and silyl protecting groups [cf T.W. Greene and P.G.M. Wuts *Protective Groups in Organic Synthesis*, J. Wiley, NY 1991]. When an aqueous solution of tetra-*n*-butylammonium hydroxide is titrated with hydrofluoric acid, a homogeneously melting octadecahydrate (m 37<sup>o</sup>) is obtained [Fowler et al. *J Am Chem Soc* **62** 1140 1940]. When the octadeca- or tri- hydrates are heated under high vacuum at temperatures of above  $80^\circ$ , decomposition occurs giving tetrabutylammonium bifluoride, tributylamine and 1-butene; with loss of a lot of its useful properties. When the trihydrate is warmed at  $40^\circ$  under high vacuum, almost “anhydrous” TBAF is obtained. This salt contained 0.1-0.3 molar equivalents of  $\text{H}_2\text{O}$  (by  $^1\text{H}$  NMR). It is an oil at room temperature [Cox et al. *J Org Chem* **49** 3216 1984]. Virtually anhydrous TBAF in DMF can be prepared without heating by using aliquots (~10 to 30mg) of the trihydrate and placing them under high vacuum (<0.1mm Hg) for 0.5 hour at  $\sim 25^\circ$ , the flask is flushed with  $\text{N}_2$ , and dry DMF (2ml) is added. The solution is transferred (syringe) to a round-bottomed flask (under  $\text{N}_2$ ), containing 4Å Molecular sieves (which had been activated by flame-drying for 5 minutes under high vacuum), and stirred with a micro stirrer bar for 30 minutes [Majetich et al. *J Org Chem* **51** 1745 1986]. This is then used for catalytic or other reactions in anhydrous media.

TBAF is also available commercially as a 1 M solution in THF, 75% solution in  $\text{H}_2\text{O}$ , on alumina support 15wt% loading [429-41-4], on silica gel at  $\sim 1.5\text{mmol/g F}^-$  as a non-hygroscopic support for use as a source of “naked” fluoride ions [Ricci et al. *Tetrahedron Lett* **23** 577 1982, Gambacorta et al. *Synth Commun* **19** 2441 1989]. [Beilstein **4** III 292, For further applications see selected volumes of Fieser & Fieser's *Reagents for Organic Synthesis*.]

## CATALYSTS—Part 2

## LIGANDS &amp; REAGENTS USED FOR MAKING LIGANDS THAT ASSIST CATALYSIS

*Ligands and reagents in this section are mainly ones that are used within this and the former section Catalysts-Part 1. Others will be found scattered in Chapters 4, 5 and 7, and can be identified by their CAS Registry Numbers through the index, or from their commonly used names.*

**( $\eta^3$ -Allyl)( $\eta^5$ -cyclopentadienyl)palladium(II) [(allyl)(cyclopentadienyl)palladium(II)] [1271-03-0] M 212.5, m 61° (dec).** This complex is volatile and should be handled in an efficient fume hood. Using Schlenk equipment under N<sub>2</sub> or argon and strictly dry conditions, a clear yellow solution of bis( $\eta^3$ -allyl)di- $\mu$ -chloro-dipalladium(II) (9.9g, 27mmol, see [12012-95-2]) in THF (100ml) and \*C<sub>6</sub>H<sub>6</sub> (100ml) is prepared, and cooled in an ice-NaCl bath to -20°. A solution of sodium cyclopentadienyl (54mmol in 28ml of THF, [4984-82-1]) in a N<sub>2</sub> flushed syringe is added dropwise to the yellow solution at -20° with stirring, whereby the colour changes to dark red. The ice bath is removed after stirring for 1 hour, the temperature is allowed to rise to ~25°, stirring is continued for 30 minutes and the solvents are removed by evaporation *in vacuo* (30-60 torr; no higher than 20° because the Pd complex will begin to sublime at ~25°) to yield a dark red solid. The residue is extracted with hexane (80ml), the extract is filtered under N<sub>2</sub> (use fluted filter paper as a glass frit is likely to become clogged). The filtrate is evaporated as before (*in vacuo* at 30-60 torr) to give red needles of the palladium(II) complex (9.2g, 80%). Note that by using mechanical stirring the yield can be improved to 98%. This product is satisfactory for most preparations of Pd(0) complexes, but an analytical sample is readily obtained as red needle-like crystals by subliming it at 40°/30mm. It has an unpleasant odour, is relatively stable in air, decomposes gradually in air at ~25° to form a black solid which is not soluble in hexane, and is best stored below -20° under N<sub>2</sub> or argon. The <sup>1</sup>H NMR (\*C<sub>6</sub>D<sub>6</sub>) has  $\delta$  at 2.14 (d, *J* = 11Hz, 2H), 3.11 (d, *J* = 6Hz, 6H) and 4.63 (m, 1H) complex for allyl protons and 8.1 (s, 5H, cyclopentadienyl protons). [Tatsuno et al. *Inorg Synth XIX* 221 1979]. The complex is useful for preparing Pd(0) complexes [Shaw *J Chem Soc* 247 1960] by reaction with hindered alkyl phosphines [Otsuka et al. *J Am Chem Soc* 98 5850 1976], reacts with isonitriles to form Pd(CNR)<sub>2</sub> complexes [Fischer & Werner *Chem Ber* 95 703 1962, Otsuka et al. *J Am Chem Soc* 91 6994 1969], and has been used to prepare BINAP complexes such as Pd[(*R*)-BINAP]<sub>2</sub> for asymmetric catalysis [Ozawa et al. *Organomet* 12 4188 1993].

**Ammonium perrhenate (NH<sub>4</sub>ReO<sub>4</sub>) [13598-65-7] M 268.2, m 365°(dec.), d<sub>4</sub><sup>25</sup> 3.97, pK<sup>25</sup> -1.25 (for HReO<sub>4</sub>).** The higher solubility of the ammonium salt in H<sub>2</sub>O (17g/L at 0°, and 162g/L at ~50°) compared to that of the potassium salt (see [10466-65-6]) has made this salt preferable for use in the preparation of a variety of rhenium compounds. It is prepared from perrhenic acid (see [13768-11-1]) which is obtained from KReO<sub>4</sub> (10.0g, 0.036mol) using the procedure of Watt & Thompson (*Inorg Synth* 7 187 1963) whereby a solution in H<sub>2</sub>O (100ml, solubility is 14% at 100°) at 90° is passed through a column of cation exchange resin (Dowex 50-Wx2, but not x8) held at 90° (preferably coated with "Instatherm", Ace Glass Co., Inc Vineland, NJ, USA, to withstand the temperature), washed with H<sub>2</sub>O, and the combined aqueous eluates are concentrated *in vacuo* down to the critical volume of 5ml. If it is concentrated further and the colour darkens, then one drop of 30% aqueous H<sub>2</sub>O<sub>2</sub> should be added to decolourise it. The solution is cooled to ~0°, and while being stirred, a chilled mixture of 2-propanol (15ml, saturated with gaseous NH<sub>3</sub>) and Et<sub>2</sub>O (50ml) is added. The cold mixture is allowed to stand for 1 hour, and the white hexagonal plates (~90% yield) are filtered onto a sintered-glass funnel, washed three times with 2-propanol/Et<sub>2</sub>O (1:9), dried *in vacuo*, then at 110° to give NH<sub>4</sub>ReO<sub>4</sub> (8.9g, 96% yield) that is 99.4-99.95% pure by microanalysis and by passage through a cation-exchange resin in H<sub>2</sub>O and titrating with standard base. [Thompson *Inorg Synth* 8 171 1966, see also Smith & Long *J Am Chem Soc* 70 354 1948.]

**Benzotriazol-1-yloxy-tris(dimethylamino)phosphonium hexafluorophosphate (BOP.PF<sub>6</sub>, Castro's reagent) [56602-33-6] M 442.3, m >130°, 147-148°, 130°.** Castro's reagent is prepared by adding dropwise hexamethylphosphoric triamide (180g, 1 mol, see [680-31-9]) to a vigorously stirred solution of phosgene in toluene (180ml, 20%) during 4 hours (use a bubbler to control the evolution of CO<sub>2</sub>). After stirring for 16 hours

at  $-25^{\circ}$ , excess of  $\text{COCl}_2$  is removed under reduced pressure.  $\text{CH}_2\text{Cl}_2$  (300ml) is added to ensure a homogeneous solution, and solid hydroxybenzotriazole monohydrate (153g, see [123333-53-9]) is added. The mixture is cooled to  $-5^{\circ}$  ( $\text{Me}_2\text{CO}/\text{Dry-ice}$ ) and  $\text{Et}_2\text{NH}$  (100ml, 1mol) is added during 15 minutes, and stirring is continued at  $0^{\circ}$  for 4 hours. The precipitated  $\text{Et}_2\text{NH}\cdot\text{HCl}$  is filtered off (sintered glass funnel), the solvent is evaporated *in vacuo*,  $\text{H}_2\text{O}$  (500ml) is added and the aqueous solution is extracted with  $\text{Et}_2\text{O}$  (3 x 100ml). This aqueous solution of BOP chloride is added to a filtered solution of  $\text{KPF}_6$  (200g) in  $\text{H}_2\text{O}$  (2L) and the precipitated  $\text{BOP}\cdot\text{PF}_6$  is extracted into  $\text{CH}_2\text{Cl}_2$ . The organic layer is dried ( $\text{MgSO}_4$ ), filtered and concentrated under reduced pressure. Addition of  $\text{Et}_2\text{O}$  causes the separation of the first crop of  $\text{BOP}\cdot\text{PF}_6$  which is collected, and further crops are obtained by addition of  $\text{CH}_2\text{Cl}_2$  to the mother liquors, concentrating and adding  $\text{Et}_2\text{O}$ . The combined crops are washed with  $\text{Et}_2\text{O}$  and dried *in vacuo* to give the salt (354g, m 147-149 $^{\circ}$ , 80%) which may be recrystallised further from  $\text{Me}_2\text{CO}/\text{Et}_2\text{O}$ . The reagent is usually pure enough for efficient coupling of peptides without this recrystallisation. Its IR(KBr) has  $\nu_{\text{max}}$  at 1010 (P-N), 840, 770, and 560 ( $\text{PF}_6^-$ )  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (acetone- $d_6$ ) has  $\delta_{\text{H}}$  at 3.0 (d, 18H,  $\text{N}(\text{CH}_3)_2$ ,  $J_{\text{H-P}} = 10\text{Hz}$ ), 7.9 (m,  $4\text{H}_{\text{arom}}$ ); and the  $^{31}\text{P}$  NMR ( $\text{CH}_2\text{Cl}_2/\text{D}_3\text{PO}_4$ ) has  $\delta$  at -43.7 (s,  $\text{P}^+$ ), +144.2 (septet,  $\text{PF}_6^-$ ). [Castro et al. *Synthesis* 751 1976.]

**Benzotriazol-1-yloxy(tripyrrolidino)phosphonium hexafluorophosphate (PyBOP)** [128625-52-5] **M 520.4, m 154-156 $^{\circ}$ , 156-157 $^{\circ}$** . It can be prepared and purified in a similar manner as for BOP except that pyrrolidine replaced  $\text{Et}_2\text{NH}$  [cf Castro's reagent above, see [56602-33-6] and Dormoy & Castro *Tetrahedron Lett* 3321 1979] by recrystallisation from  $\text{Me}_2\text{CO}/\text{Et}_2\text{O}$  and drying *in vacuo*. It is non-hygroscopic and can be stored at  $-25^{\circ}$ . It has  $^{31}\text{P}$  NMR ( $\text{CDCl}_3$ ) at  $\delta$ : 31.8 (s) and -143.7 (heptet,  $J = 713\text{Hz}$ ). It is an analogue of the BOP and a very good coupling agent which does not form the carcinogenic HMPA as byproduct [Coste *Tetrahedron Lett* 31 205 1990, Seebach *Helv Chim Acta* 77 1313 1994].

**(2-Biphenyl)-di-tert-butylphosphine [JohnPhos, 2-(di-tert-butylphosphino)biphenyl]** [224311-51-7] **M 298.4, m 86-88 $^{\circ}$** . This *Buchwald ligand* is prepared in Schlenk-type equipment under argon containing Mg turnings (617mg, 25.4mmol) and a small crystal of  $\text{I}_2$  (to activate the metal) at  $-25^{\circ}$ , which is treated with a solution of 2-bromobiphenyl (5.38g, 23.1mol, see [2052-07-5]) in THF (40ml), the mixture is refluxed for 2 hours and allowed to cool to  $25^{\circ}$ . Anhydrous  $\text{Cu(I)Cl}$  (2.40g, 24.2mmol) is added, the flask is capped with a septum, purged with argon for 2 minutes and di-tert-butylchlorophosphine (5.0g, 24.2mmol; see [13716-10-4]) is injected *via* a syringe, and the mixture is refluxed for 8 hours. The mixture is cooled to  $-25^{\circ}$ , diluted with 1:1 hexanes/ $\text{Et}_2\text{O}$  (200ml) and the suspension is filtered, the solids are washed with hexanes (60ml), and the solid material is added to 1:1 hexanes/ $\text{EtOAc}$  (150ml) followed by  $\text{H}_2\text{O}$  (100ml) and 30% aqueous  $\text{NH}_4\text{OH}$  (60ml). This slurry is stirred at  $-25^{\circ}$  for 5 minutes, the layers are separated and the organic layer is washed with brine (100ml), dried ( $\text{Na}_2\text{SO}_4$ ), filtered, evaporated *in vacuo* and the residue is recrystallised from  $\text{MeOH}$  (2 crops are collected) to provide *JohnPhos* as a white solid (4.46g, 67%) with **m 86-86.5 $^{\circ}$** . The IR (film) has  $\nu_{\text{max}}$  at 2956, 1459, 1362, 1173  $\text{cm}^{-1}$ . It has  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{H}}$  at 7.95-7.85 (m, 1H), 7.40-7.21 (m, 8H), 1.15 (d,  $J = 11.6\text{Hz}$ , 18H); the  $^{31}\text{P}$  NMR (121MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at -18.7; and  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ ) at  $\delta_{\text{C}}$ : 151.4, 150.9, 143.6, 143.5, 135.6, 135.2, 135.0, 130.4, 130.1, 128.3, 127.0, 126.7, 126.5, 126.2, 126.0, 125.6, 32.7, 32.6, 30.8, 30.6 (observed complexity from P-C coupling not assigned).

This ligand with  $\text{Pd}(\text{OAc})_2$  catalyses Suzuki coupling at  $-25^{\circ}$  between aryl halides (Br and Cl) with arylboronic acids with 0.5-1.0mol%Pd in high yields [Wolfe et al. *J Am Chem Soc* 121 9550 1999, see cyclohexyl-JohnPhos below]. The ligand was used for the amination of aryl halides and aryltriflates [Wolfe et al. *J Org Chem* 65 1158 2000]. This bulky biarylphosphine ligand was also used in the Pd-catalysed Stille cross-coupling reaction [Artamkina et al. *Synlett* 2 235 2006] and in the Pd-catalysed 2,3-diarylation of  $\alpha,\alpha$ -disubstituted-3-thiophenemethanols *via* cleavage of C-H and C-C bonds [Nakano et al. *J Org Chem* 71 8309 2006].

**(2-Biphenyl)-dicyclohexylphosphine [cyclohexyl-JohnPhos, 2-(dicyclohexylphosphino)biphenyl]** [247940-06-3] **M 350.5, m 102-106 $^{\circ}$** . This *Buchwald ligand* is prepared in a way similar to the preceding JohnPhos except that 2-bromobiphenyl in THF is converted to 2-lithiobiphenyl with *n*-BuLi at  $-78^{\circ}/45$  minutes, and followed by reaction with dicyclohexylchlorophosphine at  $-78^{\circ}/15$  minutes. The desired product is recrystallised from  $\text{MeOH}$  to give cyclohexyl-JohnPhos (71%) as white crystals **m 103 $^{\circ}$** . The IR (film) has  $\nu_{\text{max}}$  at 2916, 1441, 749  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (300MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 7.62-7.51 (m, 1H), 7.40-7.10 (m, 8H), 1.95-1.45 (m, 13H), 1.35-0.95 (m, 9H); and the  $^{31}\text{P}$  NMR (121MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at -12.7; and the  $^{13}\text{C}$  NMR (75MHz,

CDCl<sub>3</sub>) has  $\delta_C$  at 150.7, 150.4, 142.9, 142.8, 134.1, 133.8, 132.8, 130.61, 130.55, 130.5, 130.2, 130.1, 128.1, 127.3, 127.14, 127.05, 126.7, 126.4, 34.7, 34.5, 30.5, 30.2, 29.3, 29.1, 27.3, 27.2, 27.1, 26.4 (observed complexity from P-C coupling not assigned).

This ligand with Pd(OAc)<sub>2</sub> allows Suzuki coupling at low catalyst loadings (0.000001-0.02mol%Pd) between aryl halides (Br and Cl) with arylboronic acids in high yields, and tolerates a wide range of functional groups and substrate combinations including sterically hindered substrates. It was the most active catalyst system in terms of temperature of reaction, turnover numbers and steric tolerance in 1999 [Wolfe et al. *J Am Chem Soc* **121** 9550 1999.] This ligand was also used for the amination of aryl halides and aryltriflates [Wolfe et al. *J Org Chem* **66** 2560 2001], and was employed in the Pd-catalysed synthesis of 1,3,5-tris(2'-aminophenyl)-benzene from *o*-aminophenylboronic acid and 1,3,5-triiodobenzene, which may be used as a three-directional core building block for potential ionic receptors [Piatek & Slomiany *Synlett* 2027 2006].

**2,2'-Bipyridine (2,2'-dipyridyl,  $\alpha,\alpha'$ -dipyridyl)** [366-18-7] **M 156.2, m 69.7°, b 272-273°/atm, pK<sup>25</sup> 4.50.**

The reaction of pyridine with Na produces a mixture of bipyridyls which can be separated by fractional distillation where the 2,2'-isomer distils at 272.5°/atm and solidifies on cooling (m 69.5°). The distilled oil can be purified further by dissolving in Et<sub>2</sub>O, adding an equal volume of petroleum ether then the 2,3'- and 3,3'-isomers are washed out with several portions of H<sub>2</sub>O. The organic layer is evaporated and the oily residue is recrystallised from aqueous EtOH. 2,2'-Bipyridyl also sublimes at 65°/0.01mm. Its solubility in H<sub>2</sub>O is 0.5%, and it is very soluble in organic solvents. Its UV spectrum has  $\lambda_{\max}$  nm( $\epsilon$ ) 233 (10,200) and 280 (13,300) for the neutral species in H<sub>2</sub>O. Unlike the other isomers it complexes with metals, e.g. it gives an intense red colour with ferrous salts. The *picrate* has **m 69.7°** (from aqueous EtOH). It is a metalloprotease inhibitor with high affinity for Fe<sup>2+</sup> containing enzymes at 10<sup>-8</sup> M. [Smith *J Am Chem Soc* **46** 414 1924; UV Krumholz *J Am Chem Soc* **73** 3487 1951, *Beilstein* **23/8** IV 28, **23/16** V 8.] [TOXIC]

**1,3-Bis(1-adamantyl)-1,3-dihydro-2H-imidazol-2-ylidene (IAd)** [131042-77-8] **M 336.3, m 240-241°.** This is a stable *N*-heterocyclic carbene (NHC) which has been prepared in 96% yield from IAdCl by de-protonation in THF with catalytic amounts of dismyl anion [CH<sub>3</sub>S(O)CH<sub>2</sub><sup>-</sup>] in the presence of 1 equivalent of NaH, or with *tert*-BuOK. In this reaction H<sub>2</sub> is liberated and NaCl is precipitated. The carbene is stable in the absence of oxygen and moisture and recrystallises from toluene to give clear, colourless rectangular prisms with a sharp melting point that is unaltered by melting and re-solidifying. The IR has (KBr) has  $\nu_{\max}$  at 2920, 1504, 1448, 1378, 1350, 1303, 1213, 1291, 833, 695 and 495 cm<sup>-1</sup>; the <sup>1</sup>H NMR (\*C<sub>6</sub>H<sub>6</sub>-d<sub>6</sub>) has  $\delta_H$  at 1.58 (s, Ad<sub>4',6',10'</sub>, 12H), 2.01 (s, Ad<sub>3',5',7'</sub>, 6H), 2.29 (s, Ad<sub>2',8',9'</sub>, 12H), 6.91 (s, 3,4-CH, 2H); the <sup>13</sup>C NMR (\*C<sub>6</sub>H<sub>6</sub>-d<sub>6</sub>) [<sup>1</sup>H] has  $\delta_C$  at 211.43 (s, C<sub>2</sub>), 113.88 (dd, <sup>1</sup>J<sub>CH</sub> = 185.6Hz, <sup>2</sup>J<sub>CH</sub> = 13.4Hz, C<sub>4',5'</sub>), 55.99 (s, Ad<sub>1'</sub>), 44.8 (tm, <sup>1</sup>J<sub>CH</sub> = 131.3Hz, Ad<sub>2',8',9'</sub>), 37.7 (tm, <sup>1</sup>J<sub>CH</sub> = 126.8Hz, Ad<sub>4',6',10'</sub>), 30.3 (dm, <sup>1</sup>J<sub>CH</sub> = 128.9Hz, Ad<sub>3',5',7'</sub>); the <sup>15</sup>N NMR (\*C<sub>6</sub>H<sub>6</sub>-d<sub>6</sub>, ref NH<sub>4</sub>NO<sub>3</sub>) has  $\delta$  -160.5, and the EI-MS has *m/z* = 336.26. The X-ray crystal structure has been determined and showed a small N-C-N angle at the carbene centre. [Arduengo et al. *J Am Chem Soc* **113** 361 1991, see also Arduengo 336.3 et al. *J Am Chem Soc* **114** 5530 1992.] It is as effective, if not better in some cases, in many of the metal mediated catalytic reactions as other NHCs (Nitrogen Heterocyclic Carbenes). For further detail see the entry on IPr.Cl [250285-32-6]. [Arduengo USPatent 5 077 414 1991, *Chem Abstr* **116** 106289 1002, Kantchev, O'Brien & Organ *Aldrichimica Acta* **39** 97 2006, Phillips, Chan & Scheidt *Aldrichimica Acta* **42** 55 2009].

**1,3-Bis(1-adamantyl)imidazolium tetrafluoroborate (IAd.BF<sub>4</sub>)** [286014-42-4] **M 424.3, m 277-282°.** This NHC (N-Heterocyclic Carbene) precursor is prepared by established procedures from 2 mols of amine, 1 mol of glyoxal and 1 mol of formaldehyde in toluene/H<sub>2</sub>O in the presence of HBF<sub>4</sub>. It is as effective, if not better in some cases, in many of the metal mediated catalytic reactions as other NHCs. For further detail see the entry on IPr.Cl [250285-32-6]. [Arduengo USPatent 5 077 414 1991, *Chem Abstr* **116** 106289 1002, Kantchev, O'Brien & Organ *Aldrichimica Acta* **39** 97 2006, Phillips, Chan & Scheidt *Aldrichimica Acta* **42** 55 2009].

**2,2'-Bis(1,3,2-benzodioxaborole) [bis(catcholato)diboron, Cat-BB-Cat]** [13826-27-2] **M 237.8, m 189-196°, 195-198°.** This borole is prepared by distilling B<sub>2</sub>Cl<sub>4</sub> (3.95mmol [Wartik et al. *J Am Chem Soc* **71** 3265 1949, Urry et al. *J Am Chem Soc* **76** 5293 1954]) and CH<sub>2</sub>Cl<sub>2</sub> (30ml) into a reaction vessel containing catechol (8.03mmol) at -196°, and then the reaction is allowed to proceed over the temperature range ~78 to 25° for 15 hours to give a clear solution (slightly less than 16mmol of HCl is liberated). CH<sub>2</sub>Cl<sub>2</sub> and HCl are distilled off

from the reaction vessel *in vacuo* leaving a white residue which is sublimed at 120-130°/10<sup>-4</sup>mm to give the pure borole in ~70% yield. A similar result is obtained by using B<sub>2</sub>[NMe<sub>2</sub>]<sub>4</sub> instead of B<sub>2</sub>Cl<sub>4</sub>. Cat-BB-Cat has one peak at -30.7ppm (from BF<sub>3</sub>OEt<sub>2</sub> external reference) in the <sup>11</sup>B NMR (19MHz, CH<sub>2</sub>Cl<sub>2</sub>), and one peak for the aromatic protons at 2.69ppm (from solvent reference) in the <sup>1</sup>H NMR (60MHz, CH<sub>2</sub>Cl<sub>2</sub>); and its IR (KBr) has  $\nu_{\max}$  at 3058vw, 1610vw, 1597w, 1481s, 1464s, 1422m, 1400s, 1368s, 1346m, 1337m, 1285w, 1271m, 1263w, 1232s, 1214sh, 1157m, 1148s, 1140s, 1132s, 1117s, 1083m, 1037w, 1000w, 940w, 930m, 901w, 892m, 861m, 851w, 811s, 766m, 754s, 740s, 732s, 554s, 475w, 457w, and 419s cm<sup>-1</sup>. Molecular weight determination in \*C<sub>6</sub>H<sub>6</sub> indicated that it is a *monomer*, and alkaline hydrolyses provided 0.96 mole of hydrogen per mole of B-B compound. [Welch & Shore *Inorg Chem* 7 225 1968.] Cat-BB-Cat is efficient in olefin diboration catalysed by base-free Pt complexes such as Pt(COD)<sub>2</sub> and Pt(norbornene)<sub>2</sub> [Iverson & Smith *Organometallics* 16 2757 1997], consequently producing useful synthons for cross-coupling reactions and related conversions to other functional groups [Miyaura & Suzuki *Chem Rev* 95 2457 1995].

**Bis[(*R,R,S*)-diazaphos-SPE] {2,2',2'',2'''-(1,2-phenylenebis(1*R*,3*R*)-tetrahydro-5,8-dioxo-1*H*-[1,2,4]-diazaphospholo[1,2-*a*]pyridazine-2,1,3(3*H*)-triylo)tetrakis(*N*-[(1*S*)-1-phenethyl]benzamide}** [851609-33-1] M 1311.40, m 183-195°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -82° (c 1, THF), and the diastereomeric bis[(*S,S,S*)-diazaphos-SPE] {2,2',2'',2'''-(1,2-phenylenebis(1*S*,3*S*)-tetrahydro-5,8-dioxo-1*H*-[1,2,4]-diazaphospholo[1,2-*a*]pyridazine-2,1,3(3*H*)-triylo)tetrakis(*N*-[(1*S*)-1-phenethyl]benzamide} [851770-14-4] M 1311.40, m 289-299°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +15° (c 1, THF). The central diazaphospholane tetracarboxylic acid in the two molecules are enantiomeric and these are converted to the phenethylamide using the same *S*-phenethylamine. Thus the two substances are diastereomeric. They are prepared by mixing the diazaphospholane tetracarboxylic acid (0.34mmol) with 5 equivalents of PyBOP [(benzotriazol-1-yloxy)tripyrrrolidinophosphonium hexafluorophosphate, m 154-156° cf [128625-52-5]] under N<sub>2</sub>, adding degassed CH<sub>2</sub>Cl<sub>2</sub> (100ml) followed by 5 equivalents of *N,N*-diisopropylethylamine and 5 equivalents of ( $\alpha$ *S*)- $\alpha$ -methylbenzylmethanamine (*S*-2-phenylethylamine), and stirring overnight. The solution is exposed to the atmosphere and washed with saturated NaHCO<sub>3</sub> (50ml), 2M HCl (50ml), saturated NaHCO<sub>3</sub> (50ml) again, then H<sub>2</sub>O (50ml). The organic layer is dried (MgSO<sub>4</sub>), filtered, and evaporated *in vacuo*. The residue is purified by flash chromatography (eluted with 2:1 CH<sub>2</sub>Cl<sub>2</sub>/EtOAc), and the separation of diastereoisomers is accomplished by liquid chromatography using a Zorbax Rx-Sil column (4.6 x 250 column). The *R,R,S*-diastereomer (21% yield) has <sup>1</sup>H NMR (500MHz, d-THF) with  $\delta_{\text{H}}$  at 1.37 (d, 6H, *J* = 6.7Hz, CH<sub>3</sub>), 1.56 (d, 6H, *J* = 6.7Hz, CH<sub>3</sub>), 2.2-2.7 (m, 6H, CH<sub>2</sub>CH<sub>2</sub>), 2.65-3.0 (m, 2H, CH<sub>2</sub>CH<sub>2</sub>), 5.06 (dq, 2H, *J* = 7.5, 6.9Hz, CHCH<sub>3</sub>), 5.41 (dq, 2H, *J* = 7.7, 6.9Hz, CHCH<sub>3</sub>), 6.39 (d, 2H, *J* = 7.3Hz), 6.5-6.9 (m, 10H), 7.06-7.36 (m, 26H), 7.51 (m, 2H), 7.60 (d, 4H, *J* = 7.4Hz), 7.97 (d, 2H, *J* = 8.3Hz, NHCH CH<sub>3</sub>), 9.23 (d, 2H, *J* = 7.8Hz, NHCH CH<sub>3</sub>); the <sup>13</sup>C NMR (125MHz, d-THF) has  $\delta_{\text{C}}$  at 23.2 (s, CH<sub>3</sub>), 23.7 (s, CH<sub>3</sub>), 29.6 (s, CH<sub>3</sub>), 30.6 (s, CH<sub>2</sub>), 50.0 (s, CHCH<sub>3</sub>), 50.3 (s, CHCH<sub>3</sub>), 56.9 (br, PCHN), 57.6 (t, br, PCHN), peaks at 125-150 have not been assigned, 166.7 (s, CO), 168.2 (s, CO), 168.4 (s, CO), 168.5 (s, CO); the <sup>31</sup>P NMR (202.4MHz, d-THF) has  $\delta_{\text{C}}$  at 11.0 (br). The NMR spectra of the *S,S,S*-diastereomer is similar but not identical. [Clark et al. *J Am Chem Soc* 127 5040 2005, Clark & Landis *J Am Chem Soc* 125 11792 2003.]

It is a diazaphospholane ligand which displays high conversion and selectivity in Rh catalysed asymmetric hydroformylation reactions [Axtel et al. *Angew Chem. Int Ed Engl* 44 5834 2005, Clark et al. *J Am Chem Soc* 127 5040 2005, US Pat 7.071.357B.]

***R,R*(-)- and *S,S*(+)- *N,N'*-Bis(3,5-di-*tert*-butylsalicylidene)-1,2-*trans*-cyclohexanediamine (Jacobsen's ligand)** [*R,R*- 135616-40-9, *S,S*- 135616-36-3] M 546.8, m 203-206°, 205-208°, 205-208°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> ±310° (c 1, CH<sub>2</sub>Cl<sub>2</sub>). If this general ligand is to be purified, dissolve it in boiling EtOH and while cooling add H<sub>2</sub>O to the yellow solution until a yellow solid separates. This is filtered off, washed with a little of 95% aqueous EtOH and dried *in vacuo*. It is the ligand for preparing Jacobsen's Catalysts [Jacobsen et al. *J Am Chem Soc* 113 7063 1991, Deng & Jacobsen *J Org Chem* 57 4320 1992].

On a large scale, the *R,R*-ligand (~60g) is dissolved in CH<sub>2</sub>Cl<sub>2</sub> (500ml), washed with H<sub>2</sub>O (2 x 300ml) and brine (100ml), dried (Na<sub>2</sub>SO<sub>4</sub>), and the solvent is removed to yield a yellow powder with m 200-203°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> ±315° (c 1, CH<sub>2</sub>Cl<sub>2</sub>), the required <sup>1</sup>H and <sup>13</sup>C NMR and IR with  $\nu_{\max}$  (KBr) at 1595, 1631, 2869, 2960 cm<sup>-1</sup>. However, if it is felt that the product is of insufficient purity then it should be crystallised in two crops from Me<sub>2</sub>CO (1:20 w/v with typical 86-93% recovery [Larrow et al. *J Org Chem* 59 1939 1994].

A **dimeric Jacobsen ligand**, formed by joining two units with a methylene bridge has been prepared and crystallised from CH<sub>2</sub>Cl<sub>2</sub>/pentane. Its **bis-Mn catalyst** exhibited improved retention in a poly-dimethylsiloxane

membrane for the asymmetric epoxidation of olefins [Janssen et al. *Tetrahedron Asymmetry* **8** 3481 1997].

**\*1,5-Bis(3',5'-dimethoxyphenyl)penta-1E,4E-dien-3-one (dm-dba)** [39777-58-7] **M 354.1, m 132-134°**. To a cooled (ice-water bath) solution of NaOH (1.8g, 2.5equivalents) in H<sub>2</sub>O (18ml) diluted with EtOH (15ml) is added slowly 3',5'-dimethoxybenzaldehyde (3g, 1equivalent) and analytical grade acetone (0.52g, 0.5equivalent) during 15 minutes, then it is stirred at 25° for 1 hour. A yellow solid separates, and after a further hour it is filtered off, then washed with Et<sub>2</sub>O (3 x 50ml) and dried *in vacuo*. Purification by flash chromatography (*ca* 150 mesh Al<sub>2</sub>O<sub>3</sub> deactivated with 6% v/w H<sub>2</sub>O prior to use) and elution with petroleum ether (b 40°—60°)/EtOAc (4/1 v/v) gives *dm-dba* as a yellow solid (2.2g, 69%) which is recrystallised by layering a concentrated solution of CD<sub>2</sub>Cl<sub>2</sub> with Et<sub>2</sub>O (CD<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O, 1/3). The IR (CH<sub>2</sub>Cl<sub>2</sub>) has  $\nu_{\max}$  at 1652m (C=O), 1621vs (C=C), 1596w (C=C aromatic), 1572w (C=C aromatic) and 989m (CH trans) cm<sup>-1</sup>; the UV has  $\lambda_{\max}$  (THF) at 239 ( $\pi$ - $\pi^*$ ) and 371 (n- $\pi^*$ ) nm; the <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 6.48 (d, 1H, <sup>3</sup>J = 15.7Hz, H-2), 7.01 (1H, d, <sup>3</sup>J = 15.7Hz, H-1), 6.71 (2H, t, <sup>3</sup>J = 6.7Hz, 7.9Hz, H-6'), 7.62 (1H, d, <sup>3</sup>J = 8.2Hz, H-4'), 3.70 (s, 12H); and the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 188.7 (4°), 160.9 (4°), 143.2 (CH), 136.5 (CH), 125.5 (CH), 106.1 (4°), 102.7 (4°) and 55.3 (CH<sub>3</sub>). [Fairlamb et al. *Org Lett* **6** 4435 2004, Beilstein **7** IV 1747.]

**2,2'-Bis(diphenylphosphino)-1,1'-binaphthyl (BINAP)** [RS 98327-87-8] **M 622.7, m 283-286°**, [*R*-(+)-76189-55-4] **m 241-242°**, [*S*-(-)-76189-56-5] **m 241-242°**, [ $\alpha$ ]<sub>D</sub><sup>20</sup> **R+ and S- 233° (c 0.3, toluene)**. It has been prepared from (±)-BINOL which is converted to (±)-2,2'-dibromo-1,1'-binaphthyl (with Ph<sub>3</sub>PBr<sub>2</sub>) then to (±)-2,2'-bis(diphenylphosphinyl)-1,1'-binaphthyl (BINAPO) [with Mg/Ph<sub>2</sub>P(O)Cl] which is resolved into *S*-(-)- and *R*-(+)- BINAPO *via* separation of the diastereoisomeric (-)-2,3-di-*O*-benzoyl-L-tartrate salts in CHCl<sub>3</sub>/EtOAc followed by treatment with 0.75N NaOH, extraction into CHCl<sub>3</sub>, and isolation. The chiral BINAPO enantiomers are reduced to the respective *S*-(-)- and *R*-(+)- BINAP with excess of Cl<sub>3</sub>SiH [see 10025-78-1] and Et<sub>3</sub>N in toluene (100-120°) followed by treatment with 30% aqueous NaOH, extraction into toluene, isolation and purification as follows. Dissolve the enantiomer in toluene, wash it with 30% aqueous NaOH, three times with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate to ~15% of its volume and add an equal volume of degassed MeOH. Collect the solid, wash it with MeOH and dry it at 80°/0.005mm for 6 hours. Recrystallise it from a 1:1 mixture of toluene/EtOH to optical purity (**m 241-242°**). Pure (99%) (*S*)-(-)-BINAP has <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>, TMS) with  $\delta_{\text{H}}$  at 6.81 (d), 6.90 (t), 7.10 (m), 7.23 (s), 7.33 (t), 7.43 (dd), 7.82 (d) and 7.89 (d) where multiplicities are not necessarily coupled signals; and the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 145.43, 145.38, 145.16, 144.93, 144.87, 137.93, 137.84, 137.77, 137.42, 137.35, 137.31, 137.24, 135.49, 135.46, 135.40, 135.38, 134.29, 134.23, 134.09, 133.94, 133.88, 133.31, 133.24, 133.12, 132.93, 132.89, 132.76, 132.63, 132.58, 130.44, 128.29, 127.99, 127.95, 127.91, 127.80, 127.57, 127.47, 127.39, 126.41 and 125.69 (40 signals of 44 carbons, i.e. 4 signals must overlap). The purity is determined by GLC analysis using an OV-101 5m capillary column at 200-280° (argon carrier gas ?). TLC analysis (Kieselgel 60 PF<sub>254</sub>, with MeOH/CHCl<sub>3</sub> 1:19) is used to identify BINAP (R<sub>f</sub> 0.83), BINAPO (R<sub>f</sub> 0.4.2) and BINAP-monooxide (R<sub>f</sub> 0.67). (±)-BINAPO [86632-33-9] crystallises from hot toluene (solubility is 3.6g/100ml at 110°) or from hexane/toluene, and analytically pure crystals have **m 304-306°** (299-300° also reported). *S*-(-)-BINAPO [94041-18-6] crystallises from a hexane/toluene mixture and has **m 261-262°**, [ $\alpha$ ]<sub>D</sub><sup>24</sup> **-396° (c 0.47, C<sub>6</sub>H<sub>6</sub>)**, [ $\alpha$ ]<sub>D</sub><sup>24</sup> **-168° (c 0.5, EtOH)**; and *R*-(+)-BINAPO [94041-16-4], similarly recrystallised, has **m 262-263°**, [ $\alpha$ ]<sub>D</sub><sup>24</sup> **+396° (c 0.5, C<sub>6</sub>H<sub>6</sub>)**. [Takaya et al. *Org Synth* **67** 20 1989, Coll Vol **VIII** 57 1993]. [Noyori & Takaya *Acc Chem Res* **23** 345 1990, Kitamura et al. *Org Synth* **71** 1 1993, Takaya et al. *Org Synth* **72** 74 1995, Kitamura et al. *J Org Chem* **57** 4053 1992.]

§ A polymer supported version of BINAP is available.

**2R,3R-(+)-2,3-Bis(diphenylphosphino)butane (R,R-CHIRAPHOS)** [74839-84-2] **2S,3S-(-)-2,3-bis(diphenylphosphino)butane (S,S-CHIRAPHOS)** [64896-28-2] **M 426.5, m 108-109°**, *R*- [ $\alpha$ ]<sub>D</sub><sup>20</sup> **(+)** and *S*- [ $\alpha$ ]<sub>D</sub><sup>20</sup> **(-)** **200° (c 1.5, CHCl<sub>3</sub>)**. For *S,S*-CHIRAPHOS: Ph<sub>2</sub>PLi [prepared *in situ* from Ph<sub>3</sub>P (95g) and Li (5g) in THF (300ml) under N<sub>2</sub> at ~55°/1 hour then at 25°/2 hours, and the PhLi formed is decomposed with *tert*-BuCl (33g) for 45 minutes; the clear orange solution is boiled for 5 minutes then cooled to -4°] solution is treated with (+)-(2R,3R)-butanediol bis(tosylate) (35g, [64896-27-1]) in dry THF (100ml) over 1 hour with stirring, the temperature is allowed to rise to 25° and stirring is continued for 30 minutes. N<sub>2</sub> purged H<sub>2</sub>O (100ml) is added, the THF layer is separated and evaporated *in vacuo* to give impure *S,S*-chiraphos (note inversion of configuration) as a colourless oil. The oil is extracted under N<sub>2</sub> with Et<sub>2</sub>O (2 x 150ml), dried (Na<sub>2</sub>SO<sub>4</sub>), and filtered (under N<sub>2</sub>) into a solution of Ni(ClO<sub>4</sub>)<sub>2</sub>·6H<sub>2</sub>O (15g) in EtOH. The Na<sub>2</sub>SO<sub>4</sub> is washed with Et<sub>2</sub>O which

is added to the Ni solution. The oily red-brown deposit of the *Ni-chiraphos* sometimes contains some yellow solid of the Ni-perchlorate complex which is discarded (pyrophoric?). To the brown-oily (or part crystalline) mixture is added NaCNS (15g) in hot EtOH and stirred vigorously until a homogenous yellow-brown crystalline  $[\text{Ni}(\text{S,S-chiraphos})_2\text{NCS}]\text{NCS}$  is obtained. It is collected, washed thoroughly with EtOH then finally Et<sub>2</sub>O.

The Ni-NCS complex (15g), suspended in 95% EtOH (150ml) under N<sub>2</sub> is stirred, brought to the boil, and NaCN (4g) in H<sub>2</sub>O (20ml) is added rapidly when the Ni complex slowly dissolves to give a clear blood-red solution which then turns cloudy beige in colour. The hot solution is stirred until all the complex which had dissolved is converted to the yellow slurry. After cooling in ice-water, the solid is collected, washed with H<sub>2</sub>O (2 x 25ml) then rapidly with ice-cold EtOH (25ml) to give impure beige coloured *S,S-chiraphos*. This solid is purified by drying at 25°, dissolving in boiling absolute EtOH (~125ml), filtering through a frit under N<sub>2</sub> and allowing to stand at ~25° for 12 hours, whereby *S,S-chiraphos* deposits as lustrous colourless plates. A second crystallisation from absolute EtOH (60ml) provides optically pure (-)-(2*S*,3*S*)-bis(diphenylphosphino)butane (**S,S-CHIRAPHOS**) (5.5g), *m* 108-109° (sealed tube under N<sub>2</sub>) as colourless plates with  $[\alpha]_{\text{D}}^{27} -211^\circ$  (c 1.5, CHCl<sub>3</sub>), which is unchanged on further recrystallisation [Fryzuk & Bosnich *J Am Chem Soc* **99** 6262 1977]. Its <sup>31</sup>P NMR (CDCl<sub>3</sub>, with external H<sub>3</sub>PO<sub>4</sub> as reference) has  $\delta$  at -10.7, i.e. upfield from H<sub>3</sub>PO<sub>4</sub> [Slack et al. *Inorg Chem* **18** 3125 1979].

**For R,R-CHIRAPHOS:** This enantiomeric ligand is obtained essentially by the same synthesis as its enantiomer above. The (-)-(2*S*,3*S*)-butanediol bis(tosylate) [74839-83-1] is prepared in 94% yield, and has *m* 63-64°,  $[\alpha]_{\text{D}}^{25} -36.8^\circ$  (c 2.1, CHCl<sub>3</sub>), and <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 1.20 (d, *J* = 6Hz; 6H, CH<sub>3</sub>), 2.50 (s; 6H, tosyl-CH<sub>3</sub>), 4.70 (q, *J* ~ 7Hz; 2H, CH), 7.70 (m; 8H, aromatic-H). This is converted to bis[(2*R*,3*R*)-2,3-bis(diphenylphosphino)butan]thiocyanato-nickel(II) thiocyanate in 56% yield after washing with EtOH and Et<sub>2</sub>O, and decomposed with NaCN to give **R,R-CHIRAPHOS** in 32% yield with *m* 106-107°,  $[\alpha]_{\text{D}}^{27} +197^\circ$  (c 1.5, CHCl<sub>3</sub>), and <sup>1</sup>H NMR (CDCl<sub>3</sub>) with  $\delta$  at 1.10 (d, *J* = 7Hz; 3H, CH<sub>3</sub>), 1.40 (d, *J* = 6Hz; 6H, CH<sub>3</sub>), 2.60 (q, *J* ~ 7Hz; 2H, CH), 7.50 (m; 20H, aromatic-H). [Köiiner & Gerber *Chem Ber* **113** 2323 1980.]

**1,4-Bis-(diphenylphosphino)butane (dppb)** [7688-25-7] *M* 426.5, *m* 135-136°, 136-137°. Recrystallise it from EtOH [Trippett *J Chem Soc* 4263 1961]. [King *J Coord Chem* **1** 62 1971, Tolman *Chem Rev* **77** 313 1977.]

**1,2-Bis-(diphenylphosphino)ethane (DIPHOS, ethylene bis(diphenylphosphine))** [1663-45-2] *M* 398.4, *m* 139-140°, 140-142°, 143-144°, *pK*<sub>Est</sub> ~4.5. Recrystallise it from aqueous EtOH or \*C<sub>6</sub>H<sub>6</sub>. The dimethiodide, when recrystallised from MeOH has *m* 305-307°, and the dioxide when recrystallised from toluene or DMF (needles), or \*C<sub>6</sub>H<sub>6</sub> (plates) has *m* 252-254° (276-278°) [Isslieb et al. *Chem Ber* **92** 3175 1959, NMR: Aquiar et al. *J Org Chem* **29** 1660 1964, Bäckvall et al. *J Org Chem* **52** 5430 1987]. [Beilstein **16** IV 958.]

**R-(+)-1,2-Bis(diphenylphosphino)propane (R-PROPHOS)** [67884-32-6] *M* 412.4, *m* 68.5° (sealed tube under N<sub>2</sub>), 71-73°,  $[\alpha]_{\text{D}}^{26} R +186.0^\circ$ , *S* +61.0 (c 1.0, Me<sub>3</sub>CO). *S*-(-)-Propane-1,2-diol bis(*p*-toluenesulfonate) (51.9g, [60434-71-1]) in dry THF (75ml) is reacted with Ph<sub>2</sub>PLi to give the *R*-diol which is converted to its Ni-perchlorate salt and then converted into its NiNCS complex (45-55g) essentially as described for *chiraphos* above. This NCS complex (22g) is decomposed with NaCN essentially as described for *chiraphos* above, and finally the crude oily *R-prophos* (note change in absolute configuration) is dissolved in absolute EtOH at 50° under N<sub>2</sub>, allowed to cool to 25°, then held at 5° for 24 hours and the *diphosphine* (10g) that separated is collected and recrystallised from absolute EtOH (100ml) to give small colourless prisms of analytically pure *R-prophos* (7.5g) whose optical rotation is unchanged by further recrystallisation. [Fryzuk & Bosnich *J Am Chem Soc* **101** 3043 1979]. Its <sup>31</sup>P NMR (CDCl<sub>3</sub>, with external H<sub>3</sub>PO<sub>4</sub> as reference) has  $\delta$  at -20.6 (d, *J*<sub>P-P</sub> = 20.6Hz), i.e. upfield, and +1.7 ppm, i.e. downfield from H<sub>3</sub>PO<sub>4</sub> [Slack et al. *Inorg Chem* **18** 3125 1979]. Pure *R-prophos* can be prepared in a similar way.

**Bis(2-hydroxyethyl)sulfide (2,2'-thiodiglycol, thiodiglycol)** [111-48-8] *M* 112.2, *m* -16°, *b* 121-121.8°/0.01mm, 130°/2mm, 136-137°/3mm, 148°/4mm, 165°/20mm 168°/20mm, 185.5°/40mm, 194°/50mm, 282°/760mm, *d*<sub>4</sub><sup>0</sup> 1.1973, *d*<sub>4</sub><sup>25</sup> 1.1793, *n*<sub>D</sub><sup>20</sup> 1.5203, *n*<sub>D</sub><sup>26</sup> 1.5146. The sulfide has been prepared on large scales for

the manufacture of “Mustard Gas”. The thioglycol has been prepared from ethylene oxide and H<sub>2</sub>S, and the crude compound prepared from chlorohydrin contains considerable quantities of dithiane and polymeric impurities. These can be removed by distillation at 150°/8mm and the polymeric material breaks down at ~160° [Masson *J Chem Soc* **49** 236 1886]. The distillate is then diluted with H<sub>2</sub>O until its boiling point at atmospheric pressure is reduced to 165°, and superheated steam is passed through it. After evaporating off the H<sub>2</sub>O, the thiodiglycol is distilled at 147°/6mm. It is soluble in H<sub>2</sub>O, lower alcohols, CHCl<sub>3</sub>, EtOAc; and at 25° its solubility (w/w) in \*C<sub>6</sub>H<sub>6</sub> is 1.07%, in absolute Et<sub>2</sub>O it is 7.09 and in ligroin it is 0.06%. In organic acids, or alone, it is stable at 180° for many hours, but when heated at 100° with 2.5 parts of 0.1N aqueous NaOH for 30 minutes sulfide ions are formed, much more so (50%) with 1N NaOH at 140°/10 hours. Pb(OAc)<sub>2</sub>, and Cu(NO<sub>3</sub>)<sub>2</sub> decompose it at 100°, but it is stable with BaO, CaO and Al<sub>2</sub>O<sub>3</sub> even at 180°/10 hours, in unsuccessful attempts to dehydrate it to vinyl sulfide. Of the several esters of aliphatic alcohols that were reported, the *diacetate* had **b 139.5°/8mm**, the *dibutyrate* had **b 172°/8mm**, and the *dicaproate* had **b 207°/7mm** [Clayton & Reid *J Am Chem Soc* **64** 908 1942]. The *bis-3,4-diphenylcarbamoyl derivative* had **m 141.4-142.5°** (from EtOH) [Beaver et al. *J Am Chem Soc* **79** 1236 1957], and the *bis-4-nitrobenzoate ester* had **m 107.7°** (from EtOH) [Major *Bull Soc Chim Fr* **41** 634 1927]. [Nenitzescu & Scarlatescu *Chem Ber* **68** 587 1935, Beilstein **1 H** 470, **1 I** 244, **1 II** 525, **1 III** 2122, **1 IV** 2437.]

**1,2-Bis[(4*S*)-4-isopropyl-2-oxazolin-2-yl]benzene [(4*S*,4'*S*)-2,2'-(phen-1,2-diyl)bis(4-isopropyl-4,5-dihydro-oxazole)]** [131380-80-8] **M 300.4, b 175°/10<sup>-2</sup> mbar, [α]<sub>589</sub><sup>25</sup> -98.8° (c 2.13, CH<sub>2</sub>Cl<sub>2</sub>), [α]<sub>546</sub><sup>25</sup> -119.4° (c 2.13, CH<sub>2</sub>Cl<sub>2</sub>), [α]<sub>365</sub><sup>25</sup> -415.2° (c 2.13, CH<sub>2</sub>Cl<sub>2</sub>)**. This ligand is prepared by melting a small amount of anhydrous ZnCl<sub>2</sub> (68mg, 0.50mmol, m 293°) in high vacuum under argon, cooling to ~25°, adding chlorobenzene followed by *o*-phthalonitrile (1.28g, 10mmol) and *S*-valinol (3.09g, 30mmol), and boiling under reflux for 24 hours. The cooled solution is extracted with H<sub>2</sub>O (2 x 20ml), and the aqueous phase is extracted with CH<sub>2</sub>Cl<sub>2</sub> (30ml). The combined organic extracts are dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated *in vacuo*, and the residual oil is purified by flash chromatography and MPLC (pentane/EtOAc 4:1); and distilled at high vacuum to provide optically (by NMR and chiral HPLC on CHIRACEL OD) and analytically pure (C, H and N) colourless *bis-oxazoline* (2.68g, 89%). It has IR(film) with  $\nu_{\max}$  at 1655 (C=N) and 1250 (C-O) cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.95 [d, *J* = 6.8Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 1.04 [d, *J* = 6.8Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 1.88 [sept, *J* = 6.8Hz, 2H, CH(CH<sub>3</sub>)<sub>2</sub>], 4.04-4.13 (m, 4H, OCH<sub>2</sub>), 4.34-4.42 (m, 2H, CHN), 7.45-7.50 (m, 2H, 4- and 5-H), 7.74-7.77 (m, 2H, 3- and 6-H) from TMS; and the <sup>13</sup>C NMR (400MHz, CDCl<sub>3</sub>) has  $\delta$  at 17.92 (q), 18.78 (q), 32.39 (d), 70.46 (t, C-5'), 72.87 (d, C-4'), 128.61 (s, C-1, C-2), 129.97 (d, C-4, C-5), 130.36 (d, C-3, C-6), 163.95 (s, C-2'). [Bolm et al. *Chem Ber* **124** 1173 1991.] Like many chiral bis(oxazolin-2-yl) ligands it complexes with metals such as Zn (see its crystalline Zn complex [131380-93-3] in Part 1 above) and Cu, being involved in a variety of metal catalysed asymmetric synthesis such as allylation, aziridination, cyclopropanation, Diels-Alder and retro Diels-Alder, Mukaiyama aldol condensation and hydrosilylation [see reviews by Ghosh et al. *Tetrahedron: Asymmetry* **9** 1 1998, Pflatz *Acta Chem Scand* **50** 189 1996, Johnson & Evans *Acc Chem Res* **33** 325 2000, Jørgensen et al. *Acc Chem Res* **32** 605 1999.]

**1,2-Bis[(4*S*)-4-isopropyl-2-oxazolin-2-yl]ethane [(4*S*,4'*S*)-2,2'-(ethane-1,2-diyl)bis(4-isopropyl-4,5-dihydro-oxazole)]** [131380-80-8] **M 252.4, b 150°/2.10<sup>-2</sup> mbar, [α]<sub>589</sub><sup>25</sup> -93.6° (c 2.28, CH<sub>2</sub>Cl<sub>2</sub>), [α]<sub>546</sub><sup>25</sup> -112.1° (c 2.28, CH<sub>2</sub>Cl<sub>2</sub>), [α]<sub>365</sub><sup>25</sup> -335.5° (c 2.28, CH<sub>2</sub>Cl<sub>2</sub>)**. This colourless bis-oxazolinylethane (1.49g, 59% yield) was prepared by the general method described in the preceding entry from anhydrous ZnCl<sub>2</sub> (68mg, 0.50mmol), 1,2-dicyanoethane (0.80g, 10mmol), and *S*-valinol (3.09g, 30mmol) in optical and analytical purity. It has IR(film) with  $\nu_{\max}$  at 1645 (C=N) cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.87 [d, *J* = 6.7Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 0.95 [d, *J* = 6.7Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 1.74 [sept, *J* = 6.7Hz, 2H, CH(CH<sub>3</sub>)<sub>2</sub>], 2.63 [s, 4H, CH<sub>2</sub>], 3.84-3.98 (m, 4H, OCH<sub>2</sub>), 4.22 (dd, *J* = 8.5, 8.5Hz, 2H, CHN) from TMS; and the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta$  at 17.68 (q), 18.45 (q), 24.40 (t), 32.28 (d), 69.90 (t, C-5'), 71.92 (d, C-4'), 166.21 (s, C-2'). [Bolm et al. *Chem Ber* **124** 1173 1991.] Like many chiral bis(oxazolin-2-yl) ligands it complexes with metals such as Zn (see the crystalline Zn complex in Part 1 above) and Cu, being involved in a variety of metal catalysed asymmetric synthesis such as allylation, aziridination, cyclopropanation, Diels-Alder and retro Diels-Alder, Mukaiyama aldol condensation and hydrosilylation [see reviews by Ghosh et al. *Tetrahedron: Asymmetry* **9** 1 1998, Pflatz *Acta Chem Scand* **50** 189 1996, Johnson & Evans *Acc Chem Res* **33** 325 2000, Jørgensen et al. *Acc Chem Res* **32** 605 1999.]

**1,3-Bis[(4*S*)-4-isopropyl-2-oxazolin-2-yl]propane** [(4*S*,4'*S*)-2,2'-(propane-1,3-diyl)bis(4-isopropyl-4,5-dihydrooxazole)] [*131380-90-6*; **M 266.4, b 120<sup>0</sup>/2.10<sup>-2</sup> mbar**,  $[\alpha]_{589}^{25} -57.4^{\circ}$  (c 3.05, CH<sub>2</sub>Cl<sub>2</sub>),  $[\alpha]_{546}^{25} -68.4^{\circ}$  (c 3.05, CH<sub>2</sub>Cl<sub>2</sub>),  $[\alpha]_{365}^{25} -204.0^{\circ}$  (c 3.05, CH<sub>2</sub>Cl<sub>2</sub>). This colourless bis-oxazolinylpropane (1.12g, 42% yield) was prepared by the general method described in the preceding entries from anhydrous ZnCl<sub>2</sub> (68mg, 0.50mmol), 1,3-dicyanopropane (0.94g, 10mmol), and *S*-valinol (3.09g, 30mmol) in optical and analytical purity. It has IR(film) with  $\nu_{\max}$  at 1645 (C=N) cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.88 [d, *J* = 6.7Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 0.95 [d, *J* = 6.7Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>], 1.74 [sept, *J* = 6.7Hz, 2H, CH(CH<sub>3</sub>)<sub>2</sub>], 2.33-2.61 [m, 6H, CH<sub>2</sub>]<sub>3</sub>, 3.85-3.98 (m, 4H, OCH<sub>2</sub>CHN), 4.18-4.26 (m, 2H, NCHCH<sub>2</sub>O) ppm from TMS; and the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta$  at 17.77 (q), 18.37 (q), 21.66 (t), 26.21 (t), 32.27 (d), 69.87 (t), 71.94 (d), 165.53 (s). [Bolm et al. *Chem Ber* **124** 1173 1991.] Like many chiral bis(oxazolin-2-yl) ligands it complexes with metals such as Zn (see the crystalline Zn complex in the "Catalyst" section above) and Cu, being involved in a variety of metal catalysed asymmetric synthesis such as allylation, aziridination, cyclopropanation, Diels-Alder and retro Diels-Alder, Mukaiyama aldol condensation and hydrosilylation [see reviews by Ghosh et al. *Tetrahedron: Asymmetry* **9** 1 1998, Pflatz *Acta Chem Scand* **50** 189 1996, Johnson & Evans *Acc Chem Res* **33** 325 2000, Jørgensen et al. *Acc Chem Res* **32** 605 1999.]

**2,2-Bis[(4*S*)-4-isopropyl-2-oxazolin-2-yl]propane** {4,5-dihydro-2-[2(4,5-dihydro-(4*S*)-4-isopropoxyloxazol-2-yl)propan-2-yl]-(4*S*)-4-isopropoxyloxazole, (4*S*,4'*S*)-2,2'-(propane-2,2-diyl)bis(4-isopropyl-4,5-dihydro-oxazole)} [*relative stereochemistry 929270-13-3; absolute stereochemistry 797757-81-4*] **M 266.4, b 95-100<sup>0</sup>/0.5 mm, d<sub>4</sub><sup>25</sup> 0.9864, n<sub>D</sub><sup>20</sup> 1.4665,  $[\alpha]_{D}^{24} -107.5^{\circ}$  (c 1, CH<sub>2</sub>Cl<sub>2</sub>). The propane ligand for metal assisted asymmetric catalysis is prepared in two steps. The *first* is the synthesis of the bis-oxazoline precursor (-)-(4*S*,4'*S*)-*N,N'*-bis(1-hydroxymethyl-2-methylpropyl),2,2-dimethylmalonamide which is prepared in a flask under nitrogen purge at 0<sup>0</sup> (ice bath) containing a mixture of *S*-valinol (5.13g, 50mmol) and dry EtNH<sub>2</sub> (17.4ml, 124mmol, distilled from CaH<sub>2</sub>) to which 2,2-dimethylmalondioyl chloride (3.3ml, 250mmol; Evans et al. *J Org Chem* **63** 4541 1998) is added dropwise during 25 minutes (temperature rising from 0<sup>0</sup> to 10<sup>0</sup>), then allowing to warm to ~25<sup>0</sup>. After stirring for 45 minutes the colourless precipitate in the mixture is dissolved by adding CH<sub>2</sub>Cl<sub>2</sub> (120ml). Aqueous N HCl (30ml) is then added, the mixture is extracted with CH<sub>2</sub>Cl<sub>2</sub> (3 x 13ml), the combined organic layers are washed with saturated aqueous NaHCO<sub>3</sub> (30ml), brine (30ml), dried (MgSO<sub>4</sub>), filtered and evaporated *in vacuo* to provide a yellow solid which is crystallised from EtOAc (~40ml) to give the malonamide (6.4g, 84% in three crops). The amide has m 98-99<sup>0</sup>,  $[\alpha]_{D}^{24} -6.0^{\circ}$  (c 0.5, CH<sub>2</sub>Cl<sub>2</sub>); it has IR(KBr) with  $\nu_{\max}$  at 3349, 3380, 2962, 2886, 1658, 1530, 1049, 1033 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.92 (d, *J* = 6.8Hz, 6H), 0.96 (d, *J* = 6.8Hz, 6H), 1.50 (s, 6H), 1.82 (oct, *J* = 6.8Hz, 2H), 2.66 (br s, 2H), 3.52 (m, 2H), 3.69-2.86 (m, 4H), 6.41 (d, *J* = 8.6Hz, 2H); and the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>) has  $\delta$  at 18.8, 19.6, 23.7, 29.1, 50.2, 57.1, 63.5, 174.5; and correct elemental analysis for C, H and N.**

The *second* step is carried out, with stirring, under N<sub>2</sub> purge in a flask containing the preceding malondiamide (5.5g, 18.4mmol), 4-dimethylaminopyridine (204mg, 1.67mmol) in dry CH<sub>2</sub>Cl<sub>2</sub> (130ml, filtered through activated Al<sub>2</sub>O<sub>3</sub>) at ~25<sup>0</sup>, and dry EtNH<sub>2</sub> (10.25ml, 73.4mmol, distilled from CaH<sub>2</sub>) is added slowly (*via* syringe) followed by tosyl chloride (7.10g, 37mmol, 2 equivalents) dissolved in dry CH<sub>2</sub>Cl<sub>2</sub> (15ml) dropwise during 30 minutes *via* a funnel which is rinsed with CH<sub>2</sub>Cl<sub>2</sub> (2.5ml), and the mixture is stirred for 27 hours at ~25<sup>0</sup>. The mixture is then treated with saturated aqueous NH<sub>4</sub>Cl (70ml), H<sub>2</sub>O (40ml), the aqueous layer is separated, extracted with CH<sub>2</sub>Cl<sub>2</sub> (3 x 55ml), the combined organic layers are dried (MgSO<sub>4</sub>), filtered and evaporated. The oily residue is mixed with hot pentane (40ml), stirred for 5 minutes, the supernatant liquid is decanted, and the procedure repeated three times. The pentane layers are combined and evaporated *in vacuo* and the oily residue (4.05g 83%) is distilled (Kügelrohr) to give the analytically pure *bis-oxazole*. It has IR(film) with  $\nu_{\max}$  at 3308, 2960, 2874, 1746, 1525, 1353, 1303, 1037, 1017, 895, 815, 714 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.95 (d, *J* = 6.8Hz, 6H), 0.91 (d, *J* = 6.8Hz, 6H), 1.51 (s, 6H), 1.88-1.73 (m, 2H), 4.06-3.93 (m, 4H), 4.26-4.15 (m, 2H); and the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>) has  $\delta$  at 17.3, 18.5, 24.4, 32.2, 38.5, 69.9, 71.5, 168.7. [Evans et al. *Org Synth* **83** 97 2006]. Like many chiral bis(oxazolin-2-yl) ligands it complexes with metals such as Zn (see the crystalline Zn complex [131380-93-3] in the "Catalyst" section above) Cu, Ir, Pd, W, being involved (with or without further ligands) in a variety of metal catalysed asymmetric synthesis such as allylation, aziridination, cyclopropanation, Diels-Alder and retro Diels-Alder, Mukaiyama aldol condensation and hydrosilylation [see reviews by Ghosh et al. *Tetrahedron: Asymmetry* **9** 1 1998, Pflatz *Acta Chem Scand* **50** 189 1996, Johnson & Evans *Acc Chem Res* **33** 325 2000, Jørgensen et al. *Acc Chem Res* **32** 605 1999.]

**1,3-Bis(2,6-isopropylphenyl)imidazolium chloride (IPr.Cl)** [250285-32-6] **M 425.1, m 278<sup>o</sup>(dec)**. IPr.Cl is a ligand precursor of an NHC (N-heterocyclic carbene — Kantchev, O'Brien & Organ *Aldrichimica Acta* **39** 97 2006, Phillips, Chan & Scheidt *Aldrichimica Acta* **42** 55 2009) which coordinates with transition metals to form soluble catalysts that promote a variety of reactions. IPr.Cl can be prepared in two steps. *1,4-Bis(2,6-diisopropylphenyl)diazabutadiene* [1,2-bis(2,6-diisopropylphenylimino)-ethane] is first prepared by dissolving 2,6-diisopropylaniline (100g, 560mmol) and glyoxal (31.5ml, 280mmol, in 40% H<sub>2</sub>O) in absolute EtOH (500ml), and adding a few drops of formic acid as catalyst. The yellow coloured solution produces a yellow precipitate after a few hours, the mixture is stirred for 2 days, and the yellow solid is collected, washed with cold MeOH, and dried *in vacuo* to give analytically pure *diazabutadiene-ethane* (81.7g, 77.5%). It has <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) with  $\delta$  at 1.28 (d,  $J = 7.6$  Hz, 24H, CH(CH<sub>3</sub>)<sub>2</sub>), 3.03 (sep,  $J = 6.4$ Hz, 4H, CH(CH<sub>3</sub>)<sub>2</sub>), 7.27 (m, 6H, CH(CH<sub>3</sub>)<sub>2</sub>-C<sub>6</sub>H<sub>3</sub>), 8.19 (s, 2H, NCH). In the second step the *diazabutadiene-ethane* (25g, 66mmol) in toluene (500ml) is treated with solid paraformaldehyde (2.0g, 66mmol) with stirring under N<sub>2</sub>, heated to 100<sup>o</sup> until clear, then cooled to 40<sup>o</sup> and HCl (16.5ml, 66mmol, 4M in dioxane) is introduced with a syringe. The colour of the mixture turns to brown and a white precipitate separates within a few hours, but stirring is continued at ~25<sup>o</sup> for 36 hours. The solid is filtered off, washed with THF, and dried *in vacuo* to give off-white IPr.Cl (13.1g, 47%). It has <sup>1</sup>H NMR (400MHz, CD<sub>2</sub>Cl<sub>2</sub>) with  $\delta$  at 1.24 (d,  $J = 7.2$  Hz, 12H, CH(CH<sub>3</sub>)<sub>2</sub>), 1.27 (d,  $J = 7.2$  Hz, 12H, CH(CH<sub>3</sub>)<sub>2</sub>), 2.42 (sep,  $J = 6.8$ Hz, 4H, CH(CH<sub>3</sub>)<sub>2</sub>), 7.18 (t,  $J = 7.2$ Hz, 2H, *p*-C<sub>6</sub>H<sub>3</sub>), 7.4 (m, 4H, *m*-C<sub>6</sub>H<sub>3</sub>), 7.80 (s, 2H, NCH), 11.0 (s, 1H, NC(HCl)); and its <sup>13</sup>C NMR (100MHz, CD<sub>2</sub>Cl<sub>2</sub>) has  $\delta_C$ {H} at 23.9 (CH(CH<sub>3</sub>)<sub>2</sub>), 26.1 (CH(CH<sub>3</sub>)<sub>2</sub>), 86.4 (NCHCHN), 125.2 (q C, phenyl), 125.8 (CH, phenyl), 128.7 (quaternary C, phenyl), 129.5 (CH, phenyl), 145.7 (CHCl).

IPr.Cl has been used, in dioxane under argon, as a ligand with Pd<sub>2</sub>(dba)<sub>3</sub> (see above) to catalyse cross-coupling reactions between arylhalides and arylmagnesium bromides in a Kumada reaction to form diaryls efficiently [Huang & Nolan *J Am Chem Soc* **121** 9889 1999]; and in the presence of a base (*tert*-BuOK, or Cs<sub>2</sub>CO<sub>3</sub>) and Pd(OAc)<sub>2</sub> {or PdCl<sub>2</sub>, [Pd(allyl)Cl]<sub>2</sub>, PdCl<sub>2</sub>(PCy<sub>3</sub>)<sub>2</sub>} and CO under pressure it catalyses the carbonylative cross-coupling of bromopyridines with arylboronic acids to provide high yields of arylpyridylketones [Maerten et al. *Tetrahedron* **63** 682 2007].

When IPr.Cl is treated with a base (e.g. Bu'OK, Bu'ONa, K<sub>2</sub>CO<sub>3</sub> or Cs<sub>2</sub>CO<sub>3</sub>), the carbene **1,3-bis(2,6-diisopropylphenyl)-1,3-dihydro-2H-imidazol-2-ylidene** [244187-81-3] **M 388.6, m 213-217<sup>o</sup>**, is formed and is stable enough to be isolated, stored and is available commercially. The carbene-carbon atom at C2 coordinates with metals; and forms C-C bonds, by C-H insertion, with acetylene, MeCN, HCCl<sub>3</sub>, PhSOCH<sub>3</sub>, and the structures of some of the products have been confirmed by X-ray analysis [Arduengo et al. *Helv Chim Acta* **82** 2348 1999].

**Bis(pinacolato)diboron** [bis(pinacolyborane), **4, 4, 4', 4', 5, 5, 5', 5'-octamethyl-2,2'-bi-1,2,3-dioxaborolane, B<sub>2</sub>pin<sub>2</sub>**] [73183-34-3] **M 253.9, m 138<sup>o</sup>, 137-140<sup>o</sup>**. This versatile borylating agent is prepared in four steps\* from BBr<sub>3</sub> by conversion to tris(dimethylamino)borane {(Me<sub>2</sub>N)<sub>3</sub>B, m -16<sup>o</sup>, b 45<sup>o</sup>/12mm, 147-148<sup>o</sup>/atm [4375-83-1]} upon treatment with Me<sub>2</sub>NH/pentane, followed by further reaction with BBr<sub>3</sub>/pentane to give bromobis(dimethylamino)borane {(Me<sub>2</sub>N)<sub>2</sub>BBr, b 20-28<sup>o</sup>/0.5mm, 56-58<sup>o</sup>/12mm [6990-27-8]}, which on reaction with Na in toluene under reflux for ~3 hours provides tetrakis(dimethylamino)diboron [(Me<sub>2</sub>N)<sub>2</sub>B-B(Me<sub>2</sub>N)<sub>2</sub>, b 55-57<sup>o</sup>/2.5mm, 92<sup>o</sup>/12mm, Brotherton et al. *J Am Chem Soc* **82** 6242, 6245 1962] in 72% yield. Then a mixture of (Me<sub>2</sub>N)<sub>2</sub>B-B(Me<sub>2</sub>N)<sub>2</sub> [53.7g, 271mmol, one signal in the <sup>1</sup>H NMR at  $\delta$  2.67 (s, 24H) in CDCl<sub>3</sub>] in dry toluene (510ml) and pinacol (64.4g, 545mmol) in dry toluene (340ml) under dry N<sub>2</sub> is stirred in an ice-water bath, and 5.4M HCl in Et<sub>2</sub>O (203ml, 1.10 mol) is added dropwise during 2 hours. A white precipitate of Me<sub>2</sub>NH.HCl separates immediately but the mixture is stirred further for 4 hours at ~25<sup>o</sup>; the precipitate is filtered off and the filtrate is evaporated to dryness. The white residue is dissolved in pentane (~700ml), washed with H<sub>2</sub>O (3 x 500ml), dried (MgSO<sub>4</sub>), filtered and evaporated to ~150ml. This is warmed to dissolve any solid and chilled in a freezer (-30<sup>o</sup>). The first crop of **B<sub>2</sub>pin<sub>2</sub>** is collected and washed with cold pentane (2 x 30ml). The mother liquor is concentrated to give further crops of **B<sub>2</sub>pin<sub>2</sub>** which are combined with the first crop, and dried for 16 hours at ~25<sup>o</sup>/0.1mm to give colourless plates, m 138<sup>o</sup>, in 79-91% yield. The diboron is air stable and is stored in a capped container. Its IR (KBr) has  $\nu_{\max}$  at 2978, 2930, 1372, 1289, 1189, 1177, 1127, 960, 850, 744, 660, 547 cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta_H$  at 1.25 (s, 24H); the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta_C$  at 83.4, 24.9; and the <sup>11</sup>B NMR (128.3MHz, toluene) has  $\delta_B$  at 30.61 (BF<sub>3</sub>, Et<sub>2</sub>O as external reference with  $\delta$  0.00). [Ishiyama et al. *Org Synth* **77** 176 2000, Nöth *Naturforsch B: Anorg Chem, Org Chem* **39B** 1463 1984.] It borates ethylenes, acetylenes, allenes, conjugated olefins, replaces the halogen atom or a

triflate group in arenes with a  $B(OR)_2$  group effectively in the presence of catalysts such as  $Pt(PPh_3)_4$ ,  $Pt(dba)_2$ , or  $PdCl_2(dppf)$ ,  $Pd(dba)_2$  [Ishiyama et al. *J Am Chem Soc* **115** 11018 1993, Lesley *Organometallics* **15** 5137 1996, Ishiyama et al. *J Chem Soc. Chem Commun* 689 1997, Iverson & Smith *Organometallics* **16** 2757 1997, Ishiyama et al. *J Chem Soc. Chem Commun* 2073 1996, Ishiyama et al. *Tetrahedron Lett* **39** 2357 1998, Ishiyama et al. *J Org Chem* **60** 7508 1995, Ishiyama et al. *Tetrahedron Lett* **38** 3447 1997].

\* Caution: All operations should be carried out under  $N_2$  in a well ventilated fume cupboard because bromoborane derivatives fume in air by hydrolysing rapidly with evolution of heat.

**Bis(trifluoromethanesulfonyl)amine** [ $NTf_2$ , **bis(trifluoromethanesulfonyl)amide**,  $(CF_3SO_2)_2NH$ ] [82113-65-3] **M 281.2, m 49-50°, 46-57°, b 90-91°/atm, has a pK of a superacid** [ $\Delta G_{acid} = 291.8$  kcal/mol].  $NTf_2$  is a very strong neutral *Brønsted acid* which is to be compared with  $CHF_2CO_2H$  ( $\Delta G_{acid} = 323.8$ ) <  $HBr$  ( $\Delta G_{acid} = 318.3$ ) <  $CF_3CO_2H$  ( $\Delta G_{acid} = 316.3$ ) and <  $CF_3SO_3H$  ( $\Delta G_{acid} = 299.5$ ) [Koppel et al. *J Am Chem Soc* **116** 3047 1994, see pK values in AcOH below]. For purification see at the end of its synthesis below.  $NTf_2$  is a white crystalline solid that should be handled in closed systems (Schlenk equipment), or under very efficient ventilation as it is quite volatile, fumes in moist air and is very corrosive. However, it is soluble in  $H_2O$ , and it is stable in aqueous solutions where it can be titrated with NaOH as a typical acid with a pKa of ~1.7. In glacial acetic acid, the pK value by measuring  $\delta(OH)$  chemical shifts in the  $^1H$  NMR spectra gives a value of 7.8 for  $NTf_2$ , as compared with  $CF_3CO_2H$  (11.4),  $HNO_3$  (10.1),  $H_2SO_4$  (7.0),  $HI$  (5.8),  $HClO_4$  (4.9) and  $CF_3SO_3H$  (4.2) measured in the same way. If these comparisons are valid then  $NTf_2$  is a remarkably strong acid. [Foropoulos & DesMarteau *Inorg Chem* **23** 3270 1984, cf also Rode et al. *Z Phys Org Chem (Leipzig)* **253** 17 1973.]

The preparation of  $NTf_2$  in several steps from  $CH_3SO_2Cl$  [Foropoulos & DesMarteau *Inorg Chem* **23** 2720 1984] has been improved from 48% to 80% overall yield [DesMarteau & Witz *J Fluorine Chem* **52** 7 1991], and is described here.

Methanesulfonyl chloride (286g, [124-63-0], b 60°/21mm, 161°/atm, redistilled from  $P_2O_5$ , Hearst & Noller *Org Synth Coll Vol IV* 571 1963) is added slowly to solid KF (170g) and stirred for 1 hour at 25°, then  $CH_3SO_2F$  is distilled off, fraction b 123-124° is collected and redistilled from  $P_2O_5$  to give pure acid fluoride (208g). Electrochemical fluorination of the methanesulfonyl fluoride in anhydrous HF at 4-5Volts and 7-9Amps continuously for 24 hours (as described by Gramstad & Haszeldine *J Chem Soc* 173 1956) gave  $CF_3SO_2F$  (~180g, [335-05-7]) is followed by isolation *via* condensation at -78° and redistillation at -21.7°.

$CF_3SO_2F$  (76g) is bubbled into semi-frozen  $NH_3$  (600ml, at ~-77°) with dry  $N_2$ , during 0.5 hours while stirring with external cooling at -78°. Excess of  $NH_3$  is allowed to evaporate under  $N_2$  flow into a fume cupboard; NaOMe (54g) in MeOH (500ml) is added to the residual slush and the mixture is heated to 60° for a few minutes, the NaF is filtered off, the filtrate is evaporated, the residue is dried in a high vacuum to give  $CF_3SO_2NHNa$  (81g, 95%, [91742-21-1]). This Na salt (81g) is refluxed with hexamethyldisilazane (645ml, 500g, HMDS see [999-97-3]) under  $N_2$  flow (oil bath temperature <145° to avoid excessive darkening) with strong mechanical stirring as the mixture thickens progressively. After release of  $NH_3$  is complete (~12 hours), HMDS is distilled off (while increasing the vacuum), and the residual moisture sensitive solid is dried under high vacuum (in the same flask) to give  $CF_3SO_2NNa(SiMe_3)$  (106g, 92%, [91742-20-0]). This salt (106g) in dry THF (370ml) is transferred to a stirrable stainless steel autoclave (~600ml) and gaseous  $CF_3SO_2F$  (67g, ~26% excess) is transferred under  $N_2$  pressure *via* a metal vacuum system, as it is not usually possible to cool the autoclave to below -50°. The autoclave is sealed and the mixture is stirred at 100° overnight, the volatile products are vented into an efficient hood, the autoclave is washed out with  $H_2O$ , the combined  $H_2O$ /THF mixture is washed with  $CH_2Cl_2$ , evaporated in a rotavap, and the residual solid is dried under high vacuum at 110° to give  $(CF_3SO_2)_2NNa$  (129g, 98%, [91742-21-1]). The Na salt (129g) and 96%  $H_2SO_4$  (150ml) are heated in a single necked flask at 60-90°/2 x 001mm and the  $(CF_3SO_2)_2NH$  ( $NTf_2$ ) (111g, 93%) is collected in an ascending tube cooled at -22°. If necessary (cf spectra below), it is sublimed twice at 60° and/or recrystallised from  $CFCl_3$  at -50°. The analytically pure crystalline white solid has m 49-50° and can be stored in a sealed container in a drybox.

$NTf_2$  has IR (gas in equilibrium with solid at 25°) with  $\nu_{max}$  at 3395 (m), 3320 (br), 1463 (m), 1440 (m), 1300 (w), 1240 (s), 1224 (s), 1138 (s), 860 (m), 643 (vw), 614 (m), 570 (vw) and 505 (w)  $cm^{-1}$ ; Raman (solid -180°) has  $\nu_{max}$  at 3205 (w), 1464 (vw), 1458 (w), 1450 (w), 1343 (w), 1263 (s), 1142 (m, 839 (w), 778 (s), 646 (w), 591 (w), 566 (w), 537 (w), 510 (w), 392 (m), 386 (m), 346 (s), 311 (s), 276 (s), 212 (w), 195 (w) and 128 (m)  $cm^{-1}$ ; the  $^1H$  NMR (60MHz, external TMS) has peaks at 10.42 (in  $Me_2CO$   $d_6$ ), 7.92 (in  $CFCl_3$ ) ppm; the  $^{19}F$  NMR (20MHz, in  $CFCl_3$  as solvent and internal standard) has a peak at -75.97 (s) ppm, i.e. is at high field

than that of  $\text{CFCl}_3$ ; the MS[EI] has  $m/z$  at 281 ( $\text{M}^+$ ), 211 ( $\text{CF}_3\text{SO}_2\text{NSO}_2^+$ ), 147 ( $\text{CF}_3\text{SO}_2\text{N}^+$ ), 133 ( $\text{CF}_3\text{SO}_2^+$ ) and 69 ( $\text{CF}_3^+$ ); and the MS[CI] has major  $m/z$  at 282 ( $\text{MH}^+$ ), 150 ( $\text{CF}_3\text{SO}_3\text{H}^+$  or  $\text{CF}_3\text{SO}_2\text{NH}_3^+$ ) and 115 ( $\text{CF}_2\text{SO}_2\text{H}^+$  or  $\text{CF}_2\text{SONH}_2^+$ ). [Foropoulos & DesMarteau *Inorg Chem* **23** 2720 1984.]

$\text{NTf}_2$  forms several derivatives including ( $\text{CF}_3\text{SO}_2$ )<sub>2</sub> $\text{NCl}$  (viscous acrid liquid, **m**  $-96^\circ$  to  $-93^\circ$ , [91742-17-5]), ( $\text{CF}_3\text{SO}_2$ )<sub>2</sub> $\text{NNO}$  (white solid, **m**  $\sim 118^\circ$ , [91742-19-7]), ( $\text{CF}_3\text{SO}_2$ )<sub>2</sub> $\text{NNO}_2$  (white crystals **dec**  $\sim 107^\circ$ , [91742-18-6]), ( $\text{CF}_3\text{SO}_2$ )<sub>2</sub> $\text{Si}(\text{Me})_3$  (low-volatile colourless liquid, [82113-18-5]), and with CsF in MeCN it gives the cesium salt ( $\text{CF}_3\text{SO}_2$ )<sub>2</sub> $\text{NCs}$  (white crystals **m**  $115^\circ$ , [91742-16-4]) [Foropoulos & DesMarteau *Inorg Chem* **23** 2720 1984]. It forms an *N*-trimethylsilyl derivative ( $\text{TMSNTf}_2$  prepared from allylTMS and  $\text{NTf}_2$ , see above) which catalyses Diels-Alder reactions between methyl acrylate and various dienes [Mathieu & Ghosez *Tetrahedron Lett* **38** 5497 1997], as well as Friedel-Crafts acylations of anisole, and allylation and bis-allylation of carbonyl derivatives [Oshii et al. *Synlett* 1145 1997]. It readily forms the silver salt,  $\text{AgNTf}_2$ , e.g. with  $\text{Ag}_2\text{CO}_3$ .

**1,3-Bis(2,4,6-trimethylphenyl)imidazolium chloride (IMes.Cl)** [141556-45-8] **M 340.9, m**  $>300^\circ$ . This NHC (N-Heterocyclic Carbene) precursor is prepared by established procedures from 2 mols of amine, 1 mol of glyoxal and one mol of formaldehyde [Arduengo US Patent 5 077 414 1991, *Chem Abstr* **116** 106289 1002, cf Hermann et al. *Chem Eur J* **2** 1627 1996] and washed with THF before use. It is useful for making  $\text{IMesCuCl}$  which in turn is used for making  $\text{IMesAuCl}$  and  $[\text{IMesPtCl}_2]$  [Furst & Casin *Chem Commun* **46** 6924 2010].

**1,3-Bis(2,4,6-trimethylphenyl)imidazol-2-ylidene (IMes)** [141556-42-5] **M 304.4, m**  $150-155^\circ(\text{dec})$ . By using Schlenk equipment and techniques under dry  $\text{N}_2$ , this NHC is prepared from the preceding  $\text{IMes.Cl}$  (10.0g, 29.3mmol) suspended in dry THF (80ml) and stirred for 15 minutes, then solid *tert*-BuOK (3.5g, 31.2mmol) is added at  $\sim 25^\circ$  in one portion, the dark grey solution is stirred for 20 minutes, and the volatiles are removed *in vacuo*. The residue is extracted into warm toluene (2 x 50ml), filtered through Celite and the solvent is removed *in vacuo* to provide small crystals of the carbene (7.55g, 84%). The crystals are recrystallised from hexane to yield colourless crystals, **m**  $150-155^\circ(\text{dec})$  which are analytically pure and suitable for X-ray structural analysis. Its  $^1\text{H}$  NMR (300MHz, THF-*d*<sub>8</sub>, TMS) has  $\delta_{\text{H}}$  at 2.02 (s, 2',6'-CH<sub>3</sub>, 12H), 2.30 (s, 4'-CH<sub>3</sub>, 6H), 6.94 (s, ArH, 4H), 7.04 (s, 4,5-CH, 2H); and its  $^{13}\text{C}$  NMR (75MHz, THF-*d*<sub>8</sub>, TMS) has  $\delta_{\text{C}}$  at 18.04 (s, 2',6'-CH<sub>3</sub>), 21.04 (s, 4'-CH<sub>3</sub>), 121.28 (s, C-4 and C-5), 129.69 (s, Mes C-3',5'), 135.73 (s, Mes C-2',6'), 137.55 (s, Mes C-4), 139.73 (s, Mes C-1), 219.69 (s, NCN carbene-C); and the  $^{15}\text{N}$  NMR (30MHz, THF-*d*<sub>8</sub>,  $\text{NH}_4^{15}\text{NO}_3$ ) has  $\delta$  at  $-178.85$ . [Arduengo et al. *J Am Chem Soc* **114** 5530 1992.] It is useful for making a variety metal complexes of ruthenium (among other metals) for catalysing olefin metathesis [Grubbs *Tetrahedron* **60** 7117 2004], and other metathesis reactions [Love et al. *Angew Chem. Int Ed* **41** 4035 2002, Schrodi & Pederson *Aldrichimica Acta* **40** 45 2007, Kanemitsu & Seeberger *Org Lett* **5** 4541 2003].

**4-Bromobiphenylboronic acid MIDA ester** {[2-(4-bromophenyl)-6-methyl-1,3,6,2-dioxazaborocane-4,8-dione], (4-bromophenyl)[*N*-[(carboxy- $\kappa\text{O}$ )methyl]-*N*-methylglycinato(2-)- $\kappa\text{N},\kappa\text{O}$ ] boron} [943552-04-3] **M 311.9, m**  $238-240^\circ$ ,  $248-253^\circ$ . This ester is obtained from 4-bromophenylboronic acid (24.99g, 124.4mmol) and MIDA (18.31g, 124.4mmol, see [4408-64-4]) to which is added a freshly prepared mixture of DMSO (6.3ml) and toluene (118.7ml) in a flask to give a white solid suspension in a clear colourless solution. The flask is fitted with a Dean-Stark trap and a reflux condenser, and the mixture is boiled under reflux with stirring (magnetic Teflon coated bar) for 6 hours during which time the mixture became clear and about 2.1ml of  $\text{H}_2\text{O}$  are collected in the trap. Heating is stopped and the reaction is allowed to cool, with stirring, to  $\sim 25^\circ$  during 1 hour; the apparatus is washed down with toluene and the combined liquids are evaporated ( $40^\circ/15\text{mm}$ ) to remove most of the toluene and give a crude tan, chunky, solid. Acetone (15ml) is added to the residue and swirled vigorously to give a white solid in a tan solution. To this mixture is added  $\text{Et}_2\text{O}$  (150ml, dried through neutral  $\text{Al}_2\text{O}_3$  columns) in small portions (25ml) with stirring each time to allow further separation of white solid, and is finally filtered onto a medium porosity glass frit. The solid is washed with  $\text{Et}_2\text{O}$  (3 x 50ml), dried by suction, and is dried further in a vacuum ( $\sim 25^\circ/1\text{mm}$ ) for 4 hours to give 4-bromophenylboronic acid MIDA ester as an air stable, free-flowing white powder (36.3g, 94%) with **m**  $238-240^\circ$ . Alternatively, the crude tan solid can be purified by silica gel chromatography to provide the white MIDA ester in almost quantitative yield. An analytical sample has IR (in  $\text{Me}_2\text{CO}$ , thin film) with  $\nu_{\text{max}}$  at 3012, 1745, 1584, 1459, 1339, 1294, 1237, 1216, 1187, 1037, 995, 867, 812, 707  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (400MHz,  $\text{CD}_3\text{CN}$ ) has  $\delta_{\text{H}}$  at 2.50 (s, 3H), 3.89 (d,  $J = 16.0\text{Hz}$ , 2H), 4.07 (d,  $J = 16.0\text{Hz}$ , 2H), 7.41 (d,  $J = 8.0\text{Hz}$ , 2H), 7.55 (d,  $J = 8.0\text{Hz}$ , 2H); the  $^{13}\text{C}$  NMR (100MHz,  $\text{CD}_3\text{CN}$ ) has  $\delta_{\text{C}}$  at

48.2, 62.5, 124.1, 131.6, 135.2, 169.2; the  $^{11}\text{B}$  NMR (100MHz,  $\text{CD}_3\text{CN}$ ) has  $\delta_{\text{B}}$  at 12.0; the LRMS ( $\text{EI}^+$ ) has  $m/z$  (relative intensity): 314.0 (97%), 313.0 (20%), 312.0 ( $\text{M}^+\text{H}$ , 100%), 311.0 (23%), 283.0 (16%), 255.9 (16%); and the HMRS ( $\text{EI}^+$ ) has found: 312.0035, and calculated for  $\text{C}_{11}\text{H}_{12}\text{BBrNO}_4$  [ $\text{M}+\text{H}$ ] $^+$ : 312.0037. [Ballmer, Gillis and Burke *Org Synth* **86** 344 2009.]

**trans-(2-Bromovinyl MIDA boronate** {**trans-2-bromovinylboronic acid MIDA ester 6-methyl-2-trans-bromovinyl-1,3,6,2-dioxazabronane-4,8-dione, [(1Z)-2-bromoethenyl]N-[(carboxy-κO)methyl]-N-methylglycinato(2-)-κN,κO]-boron**} [*trans E* 1104636-68-1; *cis Z* has 1105069-27-9] **M 261.9**. The ester is prepared using Schlenk equipment in subdued daylight under dry  $\text{N}_2$  by adding dropwise via a syringe freshly distilled (*E*)-(2-bromoethenyl)dibromoborane (21.0g, 75.9mmol, prepared from  $\text{BBr}_2$  and acetylene, Hyuga et al. *Chem Lett* 1757 1987) to a stirred mixture of MIDA (16.9g, 114mmol, 1.5 equivalents, see [4408-64-4]) and 2,6-lutidine (17.7g, 151.9mmol, 2 equivalents) in DMSO (250ml) at  $0^\circ$  over 15 minutes, and allowed to warm to  $\sim 25^\circ$  then kept at that temperature for 48 hours. The yellow mixture is treated with  $\text{H}_2\text{O}$  (300ml) and extracted with 1:1 THF:Et $_2\text{O}$  (3 x 500ml), the combined extracts are washed with brine (3 x 350ml), dried ( $\text{MgSO}_4$ ), filtered and evaporated *in vacuo*. The light yellow solid residue is purified by flash chromatography on silica gel and eluting with Et $_2\text{O}$ /petroleum ether (1:1), then EtOAc, then EtOAc/MeCN (9:1) to give the desired boronate as colourless crystals (12g, 60%). On TLC it has  $R_{\text{F}}$  0.46 (Merck silica gel plate grade 9385, 60Å, 230-400 mesh, eluting with EtOAc and visualised with  $\text{KMnO}_4$ ). [A preparation similar to the one for vinyl MIDA boronate (this chapter) using the MIDA-di Na salt gave similar yields [cf Uno, Gillis and Burke *Tetrahedron* **65** 3230 2009]. Material stored under air at room temperature for 1.5 years showed no signs of deterioration. Crystals grown by slow evaporation of an EtOAc solution at  $\sim 25^\circ$  were suitable for X-ray analysis which revealed a *trans* configuration while confirming the expected structure with a pyramidalised boron centre. Its IR (thin film) has  $\nu_{\text{max}}$  at 3006, 2962, 1755, 1589, 1451, 1338, 1286, 1196, 1152, 1118, 1080, 1025, 1009, 961, 893, 872, 773, 678  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (500MHz,  $\text{CD}_3\text{CN}$ ,  $\delta = 1.93$  centre line, TMS) has  $\delta_{\text{H}}$  at 6.69 (d,  $J = 15.0\text{Hz}$ , 1H), 6.63 (d,  $J = 14.5\text{Hz}$ , 1H), 3.97 (d,  $J = 17.0\text{Hz}$ , 2H), 3.82 (d,  $J = 17.0\text{Hz}$ , 2H), 2.80 (s, 3H); the  $^{13}\text{C}$  NMR (125MHz,  $\text{CD}_3\text{CN}$ ,  $\delta = 1.93$  centre line, TMS) has  $\delta_{\text{C}}$  at 169.0, 118.8, 62.6, 47.9; the  $^{11}\text{B}$  NMR (100MHz,  $\text{CD}_3\text{CN}$ ,  $\text{BF}_3\cdot\text{Et}_2\text{O}$  internal standard) has  $\delta$  at 10.5; and HRMS ( $\text{EI}^+$ ) found: 261.9874, calculated for  $\text{C}_7\text{H}_{10}\text{BBrNO}_4$  [ $\text{M}+\text{H}$ ] $^+$ : 261.9886. It is very useful in the reactions stated for vinyl MIDA boronate and for iterative cross-coupling. [Lee, Gray, Peak and Burke *J Am Chem Soc* **130** 466 2008, Uno, Gillis and Burke *Tetrahedron* **65** 3230 2009, Gillis & Burke *Aldrichimica Acta* **42** 17 2009].

**Chloro(dimethylsulfide)gold(I)** [(dimethylsulfide)gold(I) chloride, aurochloro dimethylsulfide, **AuCl.Me $_2$ S**] [29892-37-3] **M 295.6, m >100°(dec), 100-102°(dec), 120°(dec)**. All procedures for preparing this complex should be carried out in an efficient fume cupboard because  $\text{Me}_2\text{S}$  is **toxic** and has a **foul** odour. The gold complex has been prepared by adding  $\text{Me}_2\text{S}$  (1 mole, see [75-18-3]) slowly to a stirred solution of  $\text{NH}_4\text{AuCl}_4$  (1mole) in dilute *Aqua Regia* when the complex separates and is collected [Allen & Wilkinson *Spectrochimica Acta* **28B** 2257 1972]. Alternatively,  $\text{Me}_2\text{S}$  is added to a solution of auric chloride ( $\text{AuCl}_3$ ) in HCl [which forms *aurichloric acid* (*chloroauric acid*)  $\text{HAuCl}_4$  see [16903-35-8]] when much heat is evolved and a flocculent white precipitate of *AuCl.Me $_2$ S* separates. It is filtered off, washed with  $\text{H}_2\text{O}$  (in which it is insoluble) and a little EtOH (in which it is slightly soluble), and dried at room temperature in air in the dark. It is air stable and can be recrystallised from  $^*\text{C}_6\text{H}_6/\text{Me}_2\text{CO}$ . Clearly, the  $\text{Me}_2\text{S}$  reduces Au(III) to Au(I) before the complex separates. It is rapidly decomposed by sunlight to give metallic gold,  $\text{Me}_2\text{S}$  and HCl. In the absence of sunlight it can be kept in the solution in which it has been prepared for long periods, or in the presence of a slight excess of  $\text{Me}_2\text{S}$  in which colourless crystalline needles are formed. It is soluble in many organic solvents, e.g.  $\text{CH}_2\text{Cl}_2$ ,  $^*\text{C}_6\text{H}_6$ , but solutions deposit metallic gold in time. The gold content can be determined by exposing solutions to sunlight, and the gold that deposits can be collected and weighed. On heating in  $\text{N}_2$ ,  $\text{Me}_2\text{S}$  begins to evolve above  $100^\circ$  until  $200^\circ$ , after which a residue of gold is obtained. [Phillips *J Am Chem Soc* **23** 250 1901.] In a third synthesis, a suspension of freshly prepared Au(I)Cl (0.23g, 1mmol, [10294-29-8]) in dry  $^*\text{C}_6\text{H}_6$  (25ml) is treated with an excess of  $\text{Me}_2\text{S}$  which causes the solid to dissolve on stirring at  $20^\circ$ . After half an hour, the solution is evaporated *in vacuo*, and the crystalline residue is recrystallised from  $^*\text{C}_6\text{H}_6/\text{Me}_2\text{CO}$  (1:1) to give a 70% yield of the pure complex, **m 100-102°(dec)**. [Dash & Schmidbaur *Chem Ber* **106** 1221 1973.]

For the **Au(III) complex** aurochloric acid (10g) in  $\text{H}_2\text{O}$  (5ml) and a 1:1 v/v mixture of  $\text{HNO}_3$  (d 1.4) and HCl (d 1.2) (1ml) are cooled in an ice bath for 10 minutes, and then  $\text{Me}_2\text{S}$  (0.5g) in  $\text{Me}_2\text{CO}$  (5ml) is added gradually (1ml at a time, at 3-5 minutes interval), heat is evolved and a yellow oil separates which soon solidifies to an

orange yellow solid (some white *aurochloro dimethylsulfide* separates simultaneously). After allowing to stand in contact with a further quantity of acid mixture, all the solid is converted to *aurichloro dimethylsulfide* ( $\text{AuCl}_3 \cdot \text{Me}_2\text{S}$ ). This is filtered off, washed with a 1:1v/v mixture of N/10  $\text{HNO}_3$  and N/10  $\text{HCl}$  then  $\text{H}_2\text{O}$  and finally with  $\text{EtOH}$ , and dried in air. It is recrystallised from  $\text{CHCl}_3/\text{Et}_2\text{O}$  to give yellow prismatic crystals of analytically pure *auri complex*  $\text{AuCl}_3 \cdot \text{Me}_2\text{S}$  with **M 365.5** [29826-91-3] and **m 160°**, which is soluble in  $\text{CHCl}_3$ ,  $\text{Et}_2\text{O}$ ,  $\text{Me}_2\text{CO}$ , and warm  $^*\text{C}_6\text{H}_6$ . When this *auri complex* is warmed with  $\text{EtOH}$  on a water bath, it is converted into the *auro complex*  $\text{AuCl} \cdot \text{Me}_2\text{S}$  and the colour changes to white. The latter can be converted back to the former complex by treatment with aqua regia or chlorine water. [Ray & Sen *J Indian Chem Soc* **7** 67 1930.]

$\text{AuCl} \cdot \text{Me}_2\text{S}$  is monomeric with  $M^+$  at *m/e* at 294; its IR (Nujol mulls between polyethylene discs or CsI plates) has bands at 730vw ( $\text{SC}_2$  *asym*), 675vw ( $\text{SC}_2$  *sym*), 345s (Au-S, str), 326s and 319sh (Au-Cl str), 279m ( $\text{SC}_2$  *def*), 198s (CSAu *def*), 109m, 93m and 83m (SAuCl *bend and lattice modes*)  $\text{cm}^{-1}$  [Goggin et al. *J Chem Soc, Dalton* 1904 1972], the  $^1\text{H}$  NMR (100MHz,  $\text{CH}_2\text{Cl}_2$ , TMS) has one peak at  $\tau = 8.1$  ppm [Allen & Wilkinson *Spectrochim Acta* **28A** 2257 1972], and the  $^1\text{H}$  NMR (60MHz,  $\text{CHCl}_3$ , TMS) has one peak at  $\delta = -2.81$  [Dash & Schmidbaur *Chem Ber* **106** 1221 1973].

It is a very useful compound for preparing a variety of gold complexes and gold catalysts [e.g. IMesAuCl in Part 1 of this Chapter].

**Chlorotris(triphenylphosphine)cobalt** [ $\text{CoCl}(\text{PPh}_3)_3$ ] [26305-75-9] **M 881.2**, **m 135-139°(dec)**, **176-179°(dec)**, **177°(dec)**, **188°(dec)**. The complex is prepared under  $\text{N}_2$ , and preferably in Schlenck-type equipment.  $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$  (9.6, 40.3mmol),  $\text{PPh}_3$  (32g, 122mmol) and  $\text{EtOH}$  (600ml) are stirred while being purged with  $\text{N}_2$  for several minutes, then vigorously at 60-70° for 30 minutes to ensure complete formation of the blue-coloured fine powder of  $\text{CoCl}_2(\text{PPh}_3)_2$ . Then under vigorous stirring the mixture is cooled to ~30° and  $\text{NaBH}_4$  (1.28g, 33.9mmol) is added during ~10 minutes (in ~10 portions) whereby the colour of the solution changes from blue to green, then to brown. The green-brown precipitate is collected by filtration (in air), washed with portions of  $\text{EtOH}$  until the filtrate is no longer blue in colour, then with  $\text{H}_2\text{O}$ ,  $\text{EtOH}$  again and finally with hexane, and dried *in vacuo* to give  $\text{CoCl}(\text{PPh}_3)_3$  (24g, 67%). In the solid state it decomposes quite slowly in air, but is stable indefinitely under argon or  $\text{N}_2$  in a refrigerator. It is soluble in  $^*\text{C}_6\text{H}_6$  and  $\text{CH}_2\text{Cl}_2$  and the spectra in  $^*\text{C}_6\text{H}_6$  have  $\lambda_{\text{max}}$  at 745, ~940sh, 1100, 2200nm. [Wakatsuki & Yamazaki *Inorg Synth* **26** 190 1989, Aresta et al. *Inorg Chim Acta* **3** 227 1969.]

$\text{CoCl}(\text{PPh}_3)_3$  is a useful stoichiometric reagent for radical dimerisation of halogenated organic compounds as in the biomimetic synthesis of the bis-sesquiterpene lactones ( $\pm$ )-biatractylolide and ( $\pm$ )-biepiasterolide [Bagal et al. *J Org Chem* **69** 9100 2004], and alkaloid dimerisation in the total synthesis of the alkaloids (-)-calycanthine and (+)-chimonanthine [Movassaghi & Schmidt *Angew Chem Int Ed* **46** 3725 2007]. It is used for the preparation of COPs (cobalt oxazoline palladacycles) catalysts [see above, Stevens & Richard *Organometallics* **18** 1346 1999].

**(1,3,5,7-Cyclooctatetraene)dilithium (2,4,6-cyclooctatriene-1,2-diyl dilithium,  $\text{Li}_2\text{C}_8\text{H}_8$ )** [40698-91-7] **M 118.1**. It is prepared by suspending lithium foil (1.0g, 144mmol) under  $\text{N}_2$  in  $\text{Et}_2\text{O}$  (20ml) and stirring at 0° with 1,3,5,7-cyclooctatetraene (5.0g, 48mmol) for 16 hours. The small amount of white precipitate is allowed to settle, and an aliquot sample of the orange solution is sucked into a syringe, and the molarity of the solution is determined by hydrolysis with  $\text{H}_2\text{O}$  and titrating with standard acid. A saturated solution of  $\text{Li}_2\text{C}_8\text{H}_8$  in  $\text{Et}_2\text{O}$  is ~0.24M. Store in the cold and away from air and moisture. The solution is not more flammable than an  $\text{Et}_2\text{O}$  solution, but the dry solid  $\text{Li}_2\text{C}_8\text{H}_8$  is **pyrophoric** in air. [Spencer *Inorg Synth* **19** 214 1979, Craswell & Spencer *Inorg Synth* **28** 127 1990.]

**(1R,2R)-(+)-1,2-Diaminocyclohexane-*N,N'*-bis(2-diphenylphosphinobenzoyl)** [(*R,R*)-DACH-phenyl Trost ligand] [138517-61-0] and **(1S,2S)-(-)-1,2-diaminocyclohexane-*N,N'*-bis(2-diphenylphosphinobenzoyl)** [(*S,S*)-DACH-phenyl Trost ligand] [169689-05-8] **M 690.8**, **m 137-142°**,  $[\alpha]_{\text{D}}^{20}$  *R,R*- +134°, *S,S*- -134° (c 1.0, MeOH),  $[\alpha]_{\text{D}}^{20}$  *R,R*- +55.1°, *S,S*- -55.1° (c 2.85,  $\text{CH}_2\text{Cl}_2$ ). The (*1R,2R*)-enantiomer is prepared from (*1R,2R*)-1,2-diaminocyclohexane (0.535g, 4.68mmol), 2-(diphenylphosphino)benzoic acid (3.02g, 9.83mmol, [17261-28-8]), DMAP (61.0g, 0.5mmol) and DCC (2.13g, 10.3mmol) in  $\text{CH}_2\text{Cl}_2$  (30ml) for 6 hours. The residue is chromatographed on silica gel and eluted with a 15—30% gradient of  $\text{EtOAc}$ /hexanes followed by recrystallisation from  $\text{CH}_2\text{Cl}_2/\text{Et}_2\text{O}$  to give the (+)-DACH-phenyl Trost ligand (2.96g, 90%). It has an  $R_{\text{F}}$  of 0.43 on silica TLC (1:1  $\text{EtOAc}$ -pentanes) and its IR (neat film from  $\text{CDCl}_3$ ) has  $\nu_{\text{max}}$  at 3303, 3070, 2935, 2857,

1955w, 18877w, 1817w, 1645s, 1538, 1478, 1434, 1328, 1306, 1162, 1091, 909 cm<sup>-1</sup>; the <sup>1</sup>H NMR (200MHz, CDCl<sub>3</sub>) has δ at 7.57 (m, 2H), 7.15-7.26 (m, 24H), 6.91 (m, 2H), 6.31 (br d, *J* = 7.7Hz, 2H, N-H), 3.77 (m, 2H), 1.87 (m, 2H), 1.62 (m, 2H), 0.9-1.3 (m 6H); the <sup>13</sup>C NMR (50MHz, CDCl<sub>3</sub>) has δ at 169.46, 140.80 (d, *J* = 24.2Hz), 137.96 (d, *J* = 11.8Hz), 137.88 (d, *J* = 12.3Hz), 136.81 (d, *J* = 21.6Hz), 134.34, 133.97 (d, *J* = 20.3Hz), 130.23, 128.79, 128.66, 128.57, 128.51, 128.43, 127.63, 127.55, 53.58, 31.71, 24.4; and has correct elemental analyses for C, H, N and P. [Trost et al. *J Am Chem Soc* **114** 9327 1992.]

The Pd complexes have found extensive use in catalysing Asymmetric Allylic Alkylation (AAA) and palladium-catalysed Dynamic Kinetic Asymmetric Transformations (DKYATs), i.e. kinetic resolutions [Trost & Fandrick *Aldrichimica Acta* **40** 59 2007].

(-)-*N,N'*-(1*R*,2*R*)-1,2-Diaminocyclohexanediylbis(2-pyridinecarboxamide) [(*R,R*)-DACH-pyridyl Trost ligand] [218290-24-5] and (+)-*N,N'*-(1*S*,2*S*)-1,2-diaminocyclohexanediylbis(2-pyridinecarboxamide) [(*R,R*)-DACH-pyridyl Trost ligand] [172138-95-3] **M** 324.4, **m** 171-176°, 174-176°, [α]<sub>D</sub><sup>20</sup> *R,R* -97.0°, *S,S* +97.0° (**c** 1.0, MeOH). The (*RS,RS*)-1,2-diaminocyclohexanediylbis(2-pyridinecarboxamide) is readily prepared by treating a mixture of 1,2-diaminocyclohexane (5.7g, 50mmol) in pyridine (20ml) and pyridine-2-carboxylic acid (12.3g, 100mmol) in pyridine (40ml) with triphenylphosphite (31.0g, 100mmol), and heating on a steam bath for 12 hours. The resulting brown oil is extracted into CHCl<sub>3</sub>, washed twice with aqueous NaHCO<sub>3</sub>, and H<sub>2</sub>O, dried (MgSO<sub>4</sub>), filtered, evaporated to a small volume and set aside overnight whereby brown crystals of the product are obtained (7.6g, 47%). Recrystallisation from CHCl<sub>3</sub> gives analytically pure (±)-ligand as white crystals, **m** 201-202°, with IR (KBr)  $\nu_{\max}$  at 3300, 3050, 2940, 2850, 1655, 1535 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>) has δ at 8.6 (m), 8.4 (br), 8.1 (m, combined 8H), 7.8 (m, 2H), 7.4 (m, 2H), 4.1 (br, 2H), 2.3 (br), 1.6 (br, combined 6H) downfield from TMS (peaks are not well separated at 60MHz). [Barnes et al. *J Chem Eng Data* **23** 349 1978.] (*1R,2R*)-DACH-pyridyl Trost ligand and [(1*S*,2*S*)-DACH-pyridyl Trost ligand are prepared similarly, but starting from (*1R,2R*)-(-)-[20439-47-8] and (*1S,2S*)(+)-[21436-03-3] 1,2-diaminocyclohexanes respectively [cf their resolution in Chapter 4, Alicyclic Compounds pp 242-243].

These ligands readily form complexes with Cu, Zn, Co, Pt and Pd [Mulqi et al. *Inorg Chim Acta* **53** L91c 1981, Adolfsen & Moberg *Tetrahedron:Asymmetry* **6** 2023 1995] and the Pd complexes have found extensive use in catalysing Asymmetric Allylic Alkylation (AAA) and palladium-catalysed Dynamic Kinetic Asymmetric Transformations (DKYATs), i.e. kinetic resolutions [Trost & Fandrick *Aldrichimica Acta* **40** 59 2007].

**Di-tert-butylchlorophosphine** [13716-10-4] **M** 180.7, **b** 48°/3mm, 72-73°/13mm, **d**<sub>4</sub><sup>25</sup> 0.951, **n**<sub>D</sub><sup>20</sup> 1.482. This chlorophosphine is prepared from two equivalents of *tert*-butylmagnesium bromide and PCl<sub>3</sub> followed by the usual workup and fractional distillation in a vacuum, and redistillation of the desired fraction. With MeLi it provides *di-tert-butylmethylphosphine* **b** 95-105°/3mm (170-172°/atm), which with MeI provides *di-tert-butyl-dimethylphosphonium iodide* **m** 95-105° (from EtOH). Hydrolysis of the chlorophosphine with H<sub>2</sub>O provides *di-tert-butylphosphine oxide* [684-19-5] **M** 162.2, **m** 55-59° (hygroscopic crystals), **b** 112°/9mm, with IR which has  $\nu_{\max}$  (CHCl<sub>3</sub>) 2950s, 2290, 1470s, 1394w, 1370, 1142s, 918, 815 (P-Bu<sup>t</sup>) and 656 cm<sup>-1</sup>; the <sup>1</sup>H NMR (CD<sub>2</sub>Cl<sub>2</sub>) has τ at 3.53 (d, *J*<sub>HP</sub> = 453Hz), 8.65 (d, *J*<sub>HCCP</sub> = 15Hz, Bu<sup>t</sup>). Note that this phosphine oxide is tautomeric with *di-tert-butylphosphinous acid*. [Hoffmann & Schellenbeck *Chem Ber* **99** 1134 1966, Hoffmann & Schellenbeck *Chem Ber* **100** 692 1967, Issleib & Krech *J Organometal Chem* **13** 283 1968, Crofts et al. *J Chem Soc C* 331 1970.]

The complex from Pd(OAc)<sub>2</sub> and this bulky *di-tert-butylchlorophosphine* ligand, after condensation with *P*-imino-azaphosphatane, catalyses efficiently the cross-coupling of arylboronic acids with aryl halides in the Suzuki-Miyaura reaction to provide specific unsymmetrical biphenyls in well over 90% yields [Kingston & Verkade *J Org Chem* **72** 2816 2007.]

**4,4'-Di-tert-butyl-2,2'-bipyridyl (dtbpy)** [72914-19-3] **M** 268.4, **b** 235°/32mm, **m** 159-160°, 159-161°, **pK**<sub>Est</sub> ~4.2. The bipyridyl has been prepared by reaction of *tert*-butylpyridine with NaNH<sub>2</sub> (ratio 2:1) in xylene at 144-218° during 26 hours, then cooled, hydrolysed with H<sub>2</sub>O, the organic layer is separated, dried, and distilled *in vacuo*. The bipyridyl solidified and can be sublimed. [McGill USP 4177349, *Chem Abstr* **92** 110871 1980, *Beilstein* **23/8** IV 181.] By coordinating with Ir(COD) and Ir(COE), complexes formed with **dtbpy** assist in C-H borylations [see Ir complexes above and Ishiyama et al. *Angew Chem. Int Edn* **41** 3056 2002, Ishiyama et al. *J Am Chem Soc* **124** 390 2002.]

**Di-tert-butylneopentylphosphine [DTBNpP, (PBU<sup>t</sup><sub>2</sub>(-CH<sub>2</sub>CMe<sub>3</sub>)]** [60633-21-8] **M 216.3, b 40°/0.1 mm, d<sup>25</sup> 0.839, n<sub>D</sub><sup>20</sup> 1.478.** This ligand, with increased steric bulk, can be obtained by adding a solution of di-tert-butylchlorophosphine (108.4g, 114.0ml, 0.6mol, see [13716-10-4]) in THF (250ml) to neopentylmagnesium chloride (from 16g, 0.67g-atom of Mg turnings, 74ml, 64g, 0.6mol of neopentyl chloride in 250ml of dry THF containing a small amount of 1,2-dibromoethane as initiator, and stirred until all the Mg has dissolved, then cooled to 0°) at 0° with stirring, allowing to warm to ~25°, boiling under reflux for *ca* 15 minutes, then the solvent is removed at ~40°/40mm). The residue is diluted with Et<sub>2</sub>O, the mixture is hydrolysed by stirring with saturated aqueous NH<sub>4</sub>Cl, the layers are separated and the organic phase is dried (MgSO<sub>4</sub>), filtered, the Et<sub>2</sub>O is removed at ~25°/50mm, and the residual phosphine is distilled at high vacuum. There are also alternative syntheses for DTBNpP. *Work with protective gloves, with eyes and face protection, and in an efficient fume cupboard.* [cf: King et al. *J Org Chem* **41** 972 1976.] It is, however, better to store and use it in the form of its salts, e.g. DTBNpP.HBF<sub>4</sub> (see below). In that case an equivalent amount of base, e.g. *t*-BuONa, must be added to the reaction mixture to liberate the free ligand. With a Pd source, e.g. Pd(OAc)<sub>2</sub> or Pd<sub>2</sub>(dba)<sub>3</sub>, H—B coupling (Hartwig—Buchwald) of various aryl halides with a variety of aromatic primary or secondary amines, or saturated heterocyclic amines, e.g. morpholine, is catalysed to form di- or tri-substituted amines in toluene solutions at 23° to 140° in high yields. [Hill et al. *J Org Chem* **71** 5117 2006.]

Unlike with Pd complexes, when DTBNpP is mixed with Pt(PhCN)<sub>2</sub>Cl<sub>2</sub> (1:1 mol per Pt atom) in CH<sub>2</sub>Cl<sub>2</sub> at 20° for a few minutes, a white binuclear complex is formed. It was characterised by spectroscopic and Xray diffraction methods to have the structure [PtCl(PBU<sup>t</sup><sub>2</sub>CH<sub>2</sub>CMe<sub>2</sub>CH<sub>2</sub>)]<sub>2</sub>, i.e. one of the methyl groups of the neopentyl moiety is folded back and is metallated by the Pt to form a five membered Pt heterocyclic ring — and the molecule is dimeric. Such metallation does not occur with PBU<sup>t</sup><sub>2</sub>Pr<sup>n</sup> under similar conditions, and in this case the monomeric *cis* and *trans* [PtCl<sub>2</sub>(PhCN)(PBU<sup>t</sup><sub>2</sub>Pr<sup>n</sup>)] square planar complexes are formed. [Mason et al. *JCS Chem Commun* 292 1976.]

**Di-tert-butylneopentylphosphonium tetrafluoroborate (DTBNpP<sup>+</sup>.HBF<sub>4</sub><sup>-</sup>)** [886059-84-3] **M 304.2, m 258-262°(dec).** This air stable salt is prepared in the same manner as [(*tert*-Bu)<sub>3</sub>PH<sup>+</sup> BF<sub>4</sub><sup>-</sup>, [131274-22-1] by mixing DTBNpP and aqueous HBF<sub>4</sub>. It is easier to handle [cf: Netherton & Fu *Org Lett* **3** 4295 2001], and is a useful ligand in Pd-catalysed cross-coupling amination reactions of aryl bromides and chlorides [Hill et al. *J Org Chem* **71** 5117 2006].

**3,3'-Di-tert-butyl-5,5',6,6'-tetramethyl-1,1'-biphenyl-2,2'-diol {H<sub>2</sub>[BIPHEN], 1,1'-biphenyl-2,2'-diol, 3,3'(1,1'-dimethylethyl)-5,5',6,6'-tetramethyl-}** [*R*- 329735-68-4, *S*- 205927-03-3, *RS*- 101203-31-0] **M 354.5, m 160.5-161.5°.** This biphenyl is prepared by the oxidation of 2-*tert*-butyl-4,5-dimethylphenol. The oxidation is carried out by slowly adding a solution of K<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> (130g, 0.442 mole, i.e. 1.02 equivalents being twice the theoretical amount to form the biphenyl) in a mixture of concentrated H<sub>2</sub>SO<sub>4</sub> (260ml) and H<sub>2</sub>O (780ml, CARE as much heat is evolved on mixing) at ~25°, to a solution of 2-*tert*-butyl-4,5-dimethylphenol (231.7g, 1.3 moles, see [1445-23-4]) in AcOH (1300ml) over 10 minutes at 55-60° and keeping at this temperature range. The colour of the solution changes from orange to green and a tan coloured solid separates during the reaction which is cooled to room temperature. It is collected, washed sequentially with H<sub>2</sub>O (2 x 250ml) and MeOH (3 x 200ml) until the solid is almost colourless, dried *in vacuo* and recrystallised from hot MeOH or dioxane/MeOH to give racemic 3,3'-di-tert-butyl-5,5',6,6'-tetramethyl-1,1'-biphenyl-2,2'-diol [101203-31-0] (~115g, ~50%) **m 160.5-161.5°** (m 140-142° before recrystallisation). Its <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>, TMS) has δ at 7.12 (s, 2H, aromatic CH), 4.79 (s, 2H, OH), 2.24 (s, 6H, CH<sub>3</sub>), 1.80 (s, 6H, CH<sub>3</sub>), 1.38 (s, 18H, C(CH<sub>3</sub>)<sub>3</sub>); and its <sup>13</sup>C NMR (100MHz, C<sub>6</sub>D<sub>6</sub>, ref C<sub>6</sub>D<sub>6</sub> at δ = 128.4) has δ at 151.24, 134.60, 134.13, 127.73, 128.79, 121.64, 35.15, 30.19, 20.46, 16.35. [Albert *J Am Chem Soc* **76** 1983 1954, Alexander et al. *J Am Chem Soc* **120** 4041 1998, Alexander et al. *Organometallics* **19** 3700 2000.]

**Optical Resolution of H<sub>2</sub>[BIPHEN]:** Two procedures have been used: (i) preparation of the respective 2,2'-cyclic phosphoric acid (*via* reaction with POCl<sub>3</sub> with the di-Na salt in THF) followed by hydrolysis and recrystallisation of the (±)-acid (-)-Cinchonidine salt in EtOH, followed by hydrolysis of the optically salts, conversion of the optically active cyclic acids to their respective optically active cyclic methyl esters, and finally cleavage with “Red-Al” [Na AlH<sub>2</sub>(OCH<sub>2</sub>CH<sub>2</sub>OMe<sub>2</sub>)] in toluene to give optically pure H<sub>2</sub>[BIPHEN]*s*. [Alexander et al. *J Am Chem Soc* **120** 4041 1998.]; and (ii) which is described here in some detail as it can be used generally for the resolution of racemic 2,2'-dihydroxybiphenyls. **Firstly:** A solution of (1*R*,2*S*,5*R*)(-)-menthol (44g, 282mmol) in CH<sub>2</sub>Cl<sub>2</sub> (100ml) is added to a solution of PCl<sub>3</sub> (58g, 423mmol, 1.5

equivalents) in  $\text{CH}_2\text{Cl}_2$  (200ml) at  $0^\circ$  during 30 minutes. The cold bath is removed, the mixture is kept at  $\sim 25^\circ$  for 1 hour, and the volatiles are removed at  $<25^\circ/100\text{mm}$ . The residual oil, which consists of *menthyl phosphorochloridite* (see [95456-31-8] in Chapter 4, p 516), is dissolved in  $\text{CH}_2\text{Cl}_2$  (250ml), and a mixture of  $\text{Et}_3\text{N}$  (118ml, 847mmol, 3 equivalents) in  $\text{CH}_2\text{Cl}_2$  (400ml) and  $\text{H}_2[\text{Biphen}]$  (100g, 282mmol) is added during 30 minutes. After 2 hours the mixture is filtered;  $\text{H}_2\text{O}_2$  (30%, 200ml, *extremely violent reaction* take great CARE, wear body protection) is added very slowly with stirring, the biphasic mixture is then stirred vigorously for 2 hours, the layers are allowed to settle, the organic phase is washed with  $\text{H}_2\text{O}$  and brine (200ml), dried ( $\text{MgSO}_4$ ), filtered and evaporated *in vacuo* to give a mixture of (*R*)- and (*S*)-  $\text{BiphenP(O)(OMen)}$  (124g, 85%) with  $^{31}\text{P}$  NMR (121MHz, THF, external ref 85%  $\text{H}_3\text{PO}_4$ ) which has  $\delta$  at -4.89 and -3.37 respectively due to the diastereomeric menthyl phosphates. **Secondly:** The diastereomeric mixture of phosphates is dissolved in the minimum volume of refluxing  $\text{AcOH}$  ( $\sim 450\text{ml}$ ), allowed to cool, and after 16 hours white crystals were filtered off, washed with cold  $\text{AcOH}$  ( $2 \times 50\text{ml}$ ) and dried *in vacuo* to give (*S*)- $\text{BiphenP(O)(OMen)}$  (42g, 97-99% de). Recrystallisation from refluxing  $\text{AcOH}$  gives pure *S*-diastereomer (37.8g, >99%de, in 61% yield), The mother liquors from the first crystallisation are evaporated to give (*R*)- $\text{BiphenP(O)(OMen)}$  (26.8g, >99%de, 43% *R*-diastereomer) after two recrystallisations from refluxing  $\text{MeOH}$  and cooling to  $0^\circ$ . [Partly racemic ester from the mother liquors are re-used in subsequent resolution processes.] **Thirdly:** To the preceding (*S*)- $\text{biphenP(O)(OMen)}$  (37.83g, 70.3mmol) in toluene (500ml) in a 2L Schlenk flask at  $0^\circ$  is treated dropwise with Red-Al (53ml, 65%wt in toluene, introduced into the separating funnel *via* a cannula) at  $0^\circ$ , and stirred for 16 hours. The mixture is carefully quenched with  $\text{H}_2\text{O}$  (75ml) then bleach (75ml), the slurry is filtered through a Celite pad, the pad is washed with toluene (250ml), the combined filtrates are allowed to settle, the toluene layer is separated, washed with bleach (200ml), brine (200ml), dried ( $\text{MgSO}_4$ ), filtered, then the toluene is removed *in vacuo* at  $0^\circ$ , any menthol (minty odour) is removed by trituration with  $\text{MeOH}$  (50ml), to provide after drying, the resolved (*S*)- $\text{H}_2[\text{BIPHEN}]$  (17.5g, 70%, >99% ee) which has  $[\alpha]_{\text{D}}^{20} -53.0^\circ$  (c 0.352, THF). The absolute configuration of this diastereoisomer was deduced from the X-ray crystallographic structure of its *syn*- $\text{Mo}(N-2,6\text{-di-iso-PrC}_6\text{H}_3(\text{CHCMe}_2\text{Ph})[(\text{S})\text{-BIPHEN}]$  complex (see Schrock-Hoveyda catalyst [205815-80-1] in Part 1). (*R*)- $\text{H}_2[\text{BIPHEN}]$  is prepared from (*R*)- $\text{BiphenP(O)(OMen)}$  using an identical procedure and the optical purity is deduced from that of its precursor as no racemisation occurs in these reactions. Note that the NMR spectra of the *RS*- (see above), *R*- and *S*- $\text{H}_2[\text{BIPHENS}]$  are the same. About 20g each of the *R*- and *S*-  $\text{H}_2[\text{BIPHENS}]$  are usually obtained from 100g of *RS*- $\text{H}_2[\text{BIPHEN}]\text{Me}$ ; and  $\sim 60\text{g}$  of the latter are pooled from the mother liquors which can be recycled. [Alexander et al. *Organometallics* **19** 3700 2000, Alexander et al. *J Am Chem Soc* **120** 4041 1998.]

By using identical compounds and reactions, but in which the *tert*-butyl group is replaced by a 1-adamantyl group the corresponding *1-adamantyl* ( $\text{H}_2\text{BIAD}$ ) derivatives are obtained in almost similar yields and used to prepare related molybdenum catalysts. For example (*S*)-3,3'-bis(1-adamantyl)-5,5',6,6'-tetramethyl-1,1'-biphenyl-2,2'-diol  $\{(\text{S})\text{-H}_2(\text{BIAD}) \{(\text{S})\text{-}(-)\text{3,3'}\text{-biphenyl-2,2'}\text{-diol, 5,5',6,6'}\text{-tetramethyl-3,3'}\text{-bis(tricyclo-}[\text{3.3.1.13.7}]\text{dec-1-yl, [255728-57-5]}\}$  with  $[\alpha]_{\text{D}}^{20} -32.1^\circ$  (c 3.3, THF), whose absolute configuration was also deduced by X-Ray crystallography of its related respective Mo complex catalyst, was prepared starting from 2-(1-adamantyl)-4,5-dimethylphenol **m 135-138**<sup>o</sup>. [Alexander et al. *Organometallics* **19** 3700 2000.]

**Dicyclohexyl(2-methylphenyl)phosphine ( $\text{Cy}_2\text{P-}o\text{-Tol}$ )** [173593-25-4] **M 288.4, m 90-93<sup>o</sup>, pK<sub>Est</sub>  $\sim 4.0$** . This phosphine is prepared and purified in the same way as  $\text{Cy}_2\text{PPh}$  (following entry), and is used in the same manner as a phosphorus ligand for catalyst systems in coupling reactions. It has the advantage that it is less susceptible to oxidation, has a higher melting point and is easier to handle.

**Dicyclohexylphenylphosphine ( $\text{Cy}_2\text{PPh}$ )** [6476-37-5] **M 274.4, m 56-57<sup>o</sup>, 57-58<sup>o</sup>, pK<sub>25</sub> 3.40**.  $\text{Cy}_2\text{PPh}$  is a "foul smelling" solid which is a good ligand for Cr, Hg and Group VIII metals. It can be prepared by reaction of  $\text{PhPCl}_2$  with  $\text{CyMgCl}$  in 50% yield [Issleib & Völker *Chem Ber* **64** 392 1961], or in 90% yield from  $\text{PhPCl}_2$  and  $\text{CyLi}$  [Screttas & Isbell *J Org Chem* **27** 2573 1973], and is recrystallised from oxygen-free  $\text{Me}_2\text{CO}$  or  $\text{EtOH}$ . In detail: to the Grignard solution, at ice-salt bath temperature, made from Mg turnings (24.1g, 0.99mol) in dry  $\text{Et}_2\text{O}$  (350ml) and  $\text{CyBr}$  (162g, 122ml, 0.99mol) in dry  $\text{Et}_2\text{O}$  (130ml), is added dropwise (2 hours) to a solution of  $\text{PhPCl}_2$  (81g, 61.4ml, 0.45mol) in dry  $\text{Et}_2\text{O}$  (100ml). The mixture is then boiled under reflux for 30 minutes and hydrolysed by stirring with saturated, deoxygenated, aqueous  $\text{NH}_4\text{Cl}$ . The  $\text{Et}_2\text{O}$  layer is separated and evaporated at  $\sim 20^\circ/30\text{mm}$  leaving an oil that crystallises after all the solvent is removed. It may be recrystallised

by slowly cooling a hot, oxygen-free EtOH solution (1 ml/g), with gentle stirring to prevent formation of an oil, to give analytically pure (Cy<sub>2</sub>PPh) as white needles (98g, 79.5%). [Bianco & Doronzo *Inorg Synth* **18** 171 1978.] It is a phosphorus ligand used in catalytic systems for coupling reactions.

**Dicyclohexylphenylphosphine oxide** (Cy<sub>2</sub>PPhO), crystallises from petroleum ether, EtOH or Me<sub>2</sub>CO and melts at **157.5°** (m 165° was also reported).

**Dicyclohexylphosphine** (Cy<sub>2</sub>PH) [829-84-5] **M 198.3, b 105-108°/3mm, 128°/8mm, 281-283°/atm, n<sub>D</sub><sup>25</sup> 1.5142, pK<sub>a</sub> 4.55.** This phosphine is a useful air sensitive ligand and should be used preferably in an inert atmosphere. It has been prepared in 80% yield by heating Cy<sub>2</sub>PSSPCy<sub>2</sub> with excess Cu and distilling [Niederfeld & Langenfeld *Chem Ber* **95** 64 1962], or in 55% yield by the addition of PH<sub>3</sub> to cyclohexene under pressure in the presence of α, α-azobis(isobutyronitrile) followed by fractional distillation [Rauhut et al. *J Org Chem* **26** 5138 1961]. It is a phosphorus ligand in catalytic systems used in coupling reactions. It forms a pale yellow *Li salt* [19966-81-5] which is slightly soluble in dioxane but insoluble in Et<sub>2</sub>O, \*C<sub>6</sub>H<sub>6</sub> and petroleum ether. [cf Edmundson *Dictionary of Organophosphorus Compounds*, Chapman & Hall, London, 1988, p 221, ISBN 0-412-25790-4.]

**Dicyclohexylphosphine oxide** (Cy<sub>2</sub>PHO) [14717-29-4] **M 214.3, m 72.5-74.5°.** The oxide crystallises from hexane or \*C<sub>6</sub>H<sub>6</sub>. It is tautomeric with dicyclohexylphosphinous acid (Cy<sub>2</sub>P-OH), whose *Ethyl ester* ([80413-46-3] **M 242.3**, obtained from CyP<sub>2</sub>Cl [16523-54-9] and EtOH) distils at **111-113°/1mm, n<sub>D</sub><sup>20</sup> 1.4950.** [Kabachnik et al. *Izv Akad Nauk SSSR, Ser Khim*, 949, 923 (*Engl Trans*) 1967]; and the *trimethylsilyl ester*, **M 286.5**, distils at **108-110°/1mm, n<sub>D</sub><sup>20</sup> 1.4919** [Foss et al. *Zh Obshch Khim* **49** 2418, 2134 (*Engl Trans*) 1979, cf Edmundson *Dictionary of Organophosphorus Compounds*, Chapman & Hall, London, 1988, p 222, ISBN 0-412-25790-4.]

**2-Dicyclohexylphosphino-2',6'-dimethoxybiphenyl (S-Phos)** [657408-07-6] **M 410.5, m 162-162.5°, 164-166°.** This Buchwald phosphine can be synthesised in one pot under argon by adding *n*-BuLi (6.20ml, 2.5M solution in hexanes, 15.5mmol, 1.1 equivalents) *via* a syringe over 5 minutes to a cold (0°) solution of 1,3-dimethoxybenzene (2.00ml, 15.3mmol, 1.1 equivalents, [151-10-0]) in dry THF (35ml), allowing the temperature to rise to ~25°, then it is stirred for 3.5 hours. The mixture is re-cooled (0°) and 2-bromochlorobenzene (1.60ml, 13.7mmol, 1.0 equivalents, [694-80-4]) is added dropwise *via* a syringe over 15 minutes while stirring vigorously; and the burgundy coloured solution is stirred for a further 15 minutes at 0°. At this stage GC analysis of an aliquot quenched with Et<sub>2</sub>O/H<sub>2</sub>O should show that all the bromochlorobenzene has been used up and a clean conversion to 2-bromo-2',6'-dimethoxybiphenyl has occurred. The mixture is cooled to -78° and *n*-BuLi (6.20ml, 2.5M solution in hexanes, 15.5mmol, 1.1 equivalents) is added dropwise *via* a syringe during 5 minutes and stirred further (with swirling if necessary) at -78° for further 30 minutes. Neat chlorodicyclohexylphosphine (3.03ml, 13.7mmol, 1.0 equivalents [16523-54-9]) is then added dropwise *via* syringe, stirred at -78° for 1 hour and allowed to warm to ~25°. The mixture is filtered through a pad of Flash silica gel topped with a layer of Celite, eluted with EtOAc (400ml) and the combined filtrates are evaporated *in vacuo* to give a yellow solid residue which is recrystallised from Me<sub>2</sub>CO to give *S-Phos* (2.90g) as a white solid. A further crop can be obtained by concentrating the mother liquors and triturating with MeOH to provide a total yield of 3.32g (59%) as a white solid m 162-162.5°. Its IR (neat) has ν<sub>max</sub> at 2921, 2848, 1588, 1470, 1459, 1430, 1245, 1109, 909, 722 cm<sup>-1</sup>; and its <sup>1</sup>H NMR (300MHz, C<sub>6</sub>D<sub>6</sub>) has δ at 7.59 (dm, *J* for the doublet 7.2Hz, 1H), 7.39-7.42 (m, 1H), 7.15-7.25 (m, 3H), 6.43 (d, *J* = 8.5 Hz, 2H), 3.33 (s, 6H), 1.60-1.94 (m, 12H), 1.06-1.36 (m, 10H) from TMS; the <sup>13</sup>C NMR (75MHz, C<sub>6</sub>D<sub>6</sub>) has δ at 158.49, 158.48, 144.3, 143.9, 137.3, 137.3, 137.0, 132.82, 132.77, 132.4, 132.2, 129.2, 128.82, 128.81, 128.2, 126.8, 121.2, 121.1, 104.0, 55.2, 35.0, 34.8, 31.0, 30.8, 30.3, 30.1, 28.22, 28.15, 28.07, 28.05, 27.40, 27.39 (observed complexity due to P-C coupling; definitive assignment have not been made); the <sup>31</sup>P NMR (121MHz, C<sub>6</sub>D<sub>6</sub>) has δ at -8.6. [Barder et al. *J Am Chem Soc* **127** 4685 2005.]

Complexation of *S-Phos* to form a Pd-catalyst [e.g. with Pd(OAc)<sub>2</sub>] provides a system with unprecedented scope, reactivity and stability for Suzuki-Miyaura coupling processes, e.g. it generates truly hindered highly chiral biaryls and heterobiaryls [Walker et al. *Angew Chem. Int Ed* **43** 1879 2004, and quinine-quinidine syntheses: Raheem et al. *J Am Chem Soc* **126** 706 2004, Huang et al. *J Am Chem Soc* **125** 6653 2003, Nguyen et al. *J Am Chem Soc* **125** 11818 2003, Gelman & Buchwald *Angew Chem. Int Ed* **42** 5993 2003], and is useful for C-N bond formation [Strieter & Buchwald *Angew Chem. Int Ed* **45** 925 2006].

**2-Dicyclohexylphosphino-2',4',6'-triisopropylbiphenyl (X-Phos, 2,4,6-triisopropyl-2'-dicyclohexylphosphinobiphenyl)** [564483-18-7] **M 476.7, m 182-184°, 187-190°**. This *Buchwald ligand* is prepared by reacting 2,4,6-triisopropylphenylmagnesium bromide with *o*-bromochlorobenzene in the presence of CuCl, and the resulting biphenyl is condensed with dicyclohexylphosphorus chloride. X-Phos is purified by solubilising it in Et<sub>2</sub>O/CH<sub>2</sub>Cl<sub>2</sub> (2:1, *ca* 30ml/g) by sonication, seeding it, and setting aside at -40° for 24 hours to give a first crop which is collected and washed with cold Et<sub>2</sub>O. A second crop can be obtained by concentrating the mother liquors then mixing with hot MeOH (~10ml, at 65°), cooling and while sonicating Et<sub>2</sub>O/CH<sub>2</sub>Cl<sub>2</sub> is added until homogeneous and set aside as before. The combined crops are dried *in vacuo* for 24 hours to give white crystals **m 182-184°**. The IR (film) has  $\nu_{\max}$  at 2923, 1461, 1449, 876 and 768 cm<sup>-1</sup>. It has <sup>1</sup>H NMR (300MHz, C<sub>6</sub>D<sub>6</sub>) with  $\delta_{\text{H}}$  at 7.48-7.42 (m, 1H), 7.24-7.06 (m, 5H), 2.85 (septet, *J* = 6.9Hz, 1H), 2.71 (septet, *J* = 6.9Hz, 2H), 1.94-1.52 (m, 12H), 1.42 (d, *J* = 6.9Hz, 6H), 1.24 (d, *J* = 6.9Hz, 6H), 1.30-1.02 (m, 10H), 1.12 (d, *J* = 6.9Hz, 6H); the <sup>31</sup>P NMR (121MHz, C<sub>6</sub>D<sub>6</sub>) has  $\delta_{\text{P}}$  at -11.5; and the <sup>13</sup>C NMR (75MHz, C<sub>6</sub>D<sub>6</sub>) has  $\delta_{\text{C}}$  at 148.7, 148.6, 148.2, 147.01, 146.99, 137.8, 137.7, 137.60, 137.558, 132.92, 132.88, 132.3, 132.2, 128.2, 127.0, 121.0, 35.4, 35.3, 35.1, 32.1, 31.9, 31.68, 31.66, 30.3, 30.1, 28.6, 28.4, 28.3, 28.2, 27.4, 26.9, 25.0, 23.92, 23.90 (observed complexity from P-C coupling). [Huang et al. *J Am Chem Soc* **125** 6653 2003, Tomori et al. *J Org Chem* **65** 5334 2000.] X-Phos has been used for the amidation of arylsulfonates, in Pd catalysed Suzuki coupling [Johnson et al. *J Org Chem* **71** 7899 2006], annulation of *o*-haloanilines [Jia & Zhu *J Org Chem* **71** 7826 2006], and the synthesis of regioregular polythiophenes [Liversidge et al. *Tetrahedron Lett* **47** 5143 2006.]

***R*(+)- and *S*(-)- *N,N*-Dimethyl-1-ferrocenylethylamine ( $\alpha$ -ferrocenylethylidimethylamine)** [*R*] 31886-58-5; [*S*] 31886-57-4; [*RS*] 31904-34-4] **M 257.2, *R* or *S* b 120-121°/0.7mm,  $d_{25}^{25}$  1.222,  $n_{\text{D}}^{20}$  1.589, *RS* b 111°/0.65mm (part dec),  $n_{\text{D}}^{25}$  1.5883, *R* [ $\alpha$ ]<sub>D</sub><sup>25</sup> + 14.1° (c 1, EtOH), *S* [ $\alpha$ ]<sub>D</sub><sup>25</sup> -14.1° (c 1.6, EtOH)**. The *RS*-amine has been prepared from ferrocenyldimethylaminoacetonitrile and MeMgI [Hauser & Lindsay *J Org Chem* **22** 906 1957], but it is easily obtained by adding a solution of 1-ferrocenylethanol (23.0g, see [1277-49-2]) in toluene (150ml) at -20° dropwise to a stirred solution of phosgene (12.5g) in toluene (100ml) at -20° during ~30 minutes, then allowing to warm to ~20°; and without isolating the chloro-derivative, the mixture is added to a solution of Me<sub>2</sub>NH (22.5g) in propan-2-ol (200ml) at ~20°. The temperature is then allowed to rise to ~25°, the solution is filtered, evaporated to dryness, treated with \*C<sub>6</sub>H<sub>6</sub> which is washed with 8.5% aqueous H<sub>3</sub>PO<sub>4</sub>, more \*C<sub>6</sub>H<sub>6</sub> is added, neutralised with Na<sub>2</sub>CO<sub>3</sub>, the \*C<sub>6</sub>H<sub>6</sub> layer is dried (K<sub>2</sub>CO<sub>3</sub>), filtered and evaporated to give the *RS*-amine (24.2g, 95% crude), which after distillation (b 110°/0.45mm with some decomposition) gives the base as a clear amber oil (17.5g, 68%). Its *picrate* [39699-90-6] forms red plates **m 136-137°** from EtOH. The crude ( $\pm$ )-base is satisfactory for optical resolution. The **resolution** is carried out by stirring a solution of the ( $\pm$ )-base (51.4g) and *R*(+)-tartaric acid (30.0g) in MeOH (100ml) at 55°, seeding, lowering the temperature slowly (2°/hour), and after 24 hours the (-)-base-(+)-acid (30.0g) is obtained which gives the (-)-base (19.0g, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -11.0°, c 1.5, EtOH). This is purified further by mixing it again with the (+)-acid (11.1g), each in MeOH (50ml) at 55°, seeded, and cooled as before to give pure (-)-base-(+)-acid from which pure (-)-base (17.0g, b 120-121°/0.7mm, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -14.1°, c 1.6, EtOH) is isolated. From the mother liquors, the crude (+)-base is converted to the (+)base-(+)tartrate, recrystallised twice from Me<sub>2</sub>CO/H<sub>2</sub>O (10:1), from which pure (+)-base, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -14.1° is obtained as a clear amber oil. [Marquarding et al. *J Am Chem Soc* **92** 3589 1970.] The absolute configuration was determined by X-Ray analysis of the (*S,R,S*)-2-(*p*-methoxyphenyl)hydroxymethyl-*N,N*-dimethyl-1-ferrocenylethylamine [yellow powder from *n*-heptane has **m 110-111°**, [ $\alpha$ ]<sub>D</sub><sup>22</sup> -120.8° (c 1.6, EtOH)] which is obtained by lithiation of the (+)-base, which is highly stereoselective, followed by reaction with *p*-anisaldehyde. [Battelle et al. *J Am Chem Soc* **95** 482 1973.]

This ferrocene is a very useful compound for preparing a variety of chiral ferrocene derivatives [see Marquarding et al. *J Chem Res (S)* 82 1977 and (*M*) 0915 1977], and remarkable phosphine ligands for catalysts by virtue of its highly stereoselective lithiation (with BuLi) in the position *ortho* to the dimethylaminoethyl group that allows substitution to give chiral 2-phosphines [Hayashi et al. *Tetrahedron Lett* 4405 1974], 2—SiMe<sub>3</sub>, 2—CH<sub>2</sub>OH and 2—CPh<sub>2</sub>OH derivatives [Marquarding et al. *J Am Chem Soc* **92** 3589 1970]. The chiral 2-lithio derivative (prepared with *sec*-BuLi/Et<sub>2</sub>O) reacts with sulfur, selenium or tellurium to give the yellow, red or red-brown **dichalcogenides** (*R,S*)-{[E-C<sub>5</sub>H<sub>3</sub>-CHMe(NMe<sub>2</sub>)]Fe(C<sub>5</sub>H<sub>5</sub>)}<sub>2</sub> (from the *R*(+)-base) and (*S,R*)-{[E-C<sub>5</sub>H<sub>3</sub>-CHMe(NMe<sub>2</sub>)]Fe(C<sub>5</sub>H<sub>5</sub>)}<sub>2</sub> (from the *R*(+)-base), where E is S, Se or Te respectively. These dichalcogenides are effective chiral ligands for the rhodium(I) {using [Rh(COD)Cl]<sub>2</sub>}—catalysed asymmetric hydrosilylation of alkyl and aryl ketones with high enantiomeric excesses. The iridium(I) catalyst exhibits lower selectivity [Nishibayashi et al. *JCS Chem Commun* 1375 1994].

The tellurium dichalcogenide (*R,S*)-{Te-C<sub>5</sub>H<sub>3</sub>-CHMe(NMe<sub>2</sub>)]Fe(C<sub>5</sub>H<sub>5</sub>)}<sub>2</sub>, [*R,S*-(*-*)-(Fc\*Te)<sub>2</sub>] is prepared from lithiated *R*-(+)-*N,N*-dimethyl-1-ferrocenylethylamine (with *sec*-BuLi/Et<sub>2</sub>O) in Et<sub>2</sub>O under N<sub>2</sub> at 0° which is then treated with Te powder in portions, and the mixture is subjected to ultrasonic irradiation for 1 hour. The mixture is poured into H<sub>2</sub>O, and air is bubbled through it for 3 hours at ~25°. The solid is collected and purified by column chromatography on activated Al<sub>2</sub>O<sub>3</sub>, and eluting with EtOAc/hexane to give analytically pure [*R,S*-(*-*)-(Fc\*Te)<sub>2</sub>] as a red-brown solid (42%) with *m* 57-58°, [α]<sub>D</sub><sup>25</sup> -622° (c 1, CHCl<sub>3</sub>), the <sup>1</sup>H NMR (270MHz, CDCl<sub>3</sub>) has δ<sub>H</sub> at 1.25 (d, *J* = 6.9Hz, 6H), 2.18 (s, 12H), 4.01 (q, *J* = 6.9Hz, 2H), 4.06 (s, 10H), 4.18 (q, *J* = 1.3Hz, 2H), 4.23 (q, *J* = 1.1Hz, 2H), 4.48 (q, *J* = 1.3Hz, 2H); the <sup>13</sup>C NMR (67.8MHz, CDCl<sub>3</sub>) has δ<sub>C</sub> at 9.6 (q), 39.7 (q), 50.2 (q), 59.5 (q), 67.1 (d), 67.7 (d), 70.7 (d), 76.2 (d), 93.7 (s). [*S,R*-(+)-(Fc\*Te)<sub>2</sub>] is prepared similarly, but from *S*-(*-*)-*N,N*-dimethyl-1-ferrocenylethylamine in 47% yield and has identical physical properties except for [α]<sub>D</sub><sup>25</sup> +613° (c 1, CHCl<sub>3</sub>). After conversion into their anions (e.g. Te<sup>-</sup>Fc\*, with NaBH<sub>4</sub>/EtOH), they react with allylic bromides, e.g. geranyl bromide, to form allylic-TeF\*c which produce the corresponding allylic alcohols (on treatment with Bu<sup>t</sup>OOH/toluene) in 14-22% enantiomeric excess by chirality transfer, possibly *via* a [2,3]-sigmatropic rearrangement [Chiba et al. *Tetrahedron Lett* **36** 1519 1995].

Similarly, but using Se instead of Te, *R*-(+)- and *S*-(*-*)-*N,N*-dimethyl-1-ferrocenylethylamine provide [*R,S*-(Fc\*Se)<sub>2</sub>] (*m* 98-100° from hexane) and [*S,R*-(Fc\*Se)<sub>2</sub>] (*m* 103° from hexane) respectively as red solids, and these catalysts induce highly enantioselective selenoxide elimination to produce axially chiral allene carboxylic esters with high enantiomeric excess [Nishibayashi et al. *Tetrahedron Lett* **35** 3115 1994].

[*R,S*-(Fc\*S)<sub>2</sub>] (*m* 169-170° from hexane) is obtained as a yellow solid in the same way, but using S instead of Se or Te. [See above for the Rh(I)-catalysed asymmetric hydrosilylation reaction of ketones with these dichalcogenide ligands.]

(1*R*,2*R*)-(+)- and (1*S*,2*S*)-(-)-1,2-Diphenylethylenediamine [(1*R*,2*R*)-(+)- and (1*S*,2*S*)-(-)-1,2-diamino-1,2-diphenylenediamine, 1,2-diamino-1,2-diphenylethane, stilbenediamine, DPEN, STEIN] [(*R*) 35132-20-8]; (*S*) 29841-69-8; (*RS*±) 16635-95-3] *M* 212.3, *R* or *S* *m* 79-83°, *RS* *m* 82°, [α]<sub>D</sub><sup>20</sup> *R,R* +102°, *S,S* -102° (c 1, EtOH), [α]<sub>D</sub><sup>23</sup> *R,R* +106°, *S,S* -106° (c 1.1, MeOH), p*K*<sub>(Est1)</sub><sup>25</sup> ~5.86, p*K*<sub>(Est2)</sub><sup>25</sup> ~8.92. The racemate is prepared in two steps. In the first, a mixture of benzil (158g, 0.75mol, [134-81-6]), anhydrous NH<sub>4</sub>OAc (400g) and cyclohexanone (80ml, 0.77mol) in glacial AcOH (1.0L) are stirred and heated under reflux (colour changing from light yellow to dark green) for 1.5 hours, and while hot it is poured into H<sub>2</sub>O (3.0L) with vigorous stirring. After cooling to ~25° overnight, the crystals are collected, washed with H<sub>2</sub>O (4 x 300ml), and crushed in a mortar (*in vacuo*) to give 2,2-spirocyclohexane-4,5-diphenyl-2H-imidazole (2,3-diphenyl-1,4-diazaspiro-[4,5]deca-1,3-diene, [5396-98-5]) (~208g, 96%) as yellow-green crystals, *m* 107-108°, on recrystallisation from hexane or aqueous MeOH. Its <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>, TMS) has δ<sub>H</sub> at 1.65-1.92 (m, 6H), 1.95-2.00 (m, 4H), 7.33-7.53 (m, 10H); and its <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has δ at 24.1, 25.7, 34.7, 104.1, 128.3, 128.9, 129.9, 133.1, 164.0. In the second step, under argon and a Dry-ice condenser, the preceding crude spiroimidazole (72.0g, 0.25mol) in dry THF (400ml, distilled from Na/benzophenone) is stirred until clear, cooled to -78° (Dry-ice/Me<sub>2</sub>CO bath), a stream of gaseous NH<sub>3</sub> is passed through until the volume of homogeneous solution increases to 400ml. Lithium metal (6.94g, 1.0mol, *via* a powder funnel, from wire cut with scissors under a gentle stream of argon) is added at such a rate that the temperature is kept below -65°. The mixture is stirred for 30 minutes, absolute EtOH (30ml, 1.0mol) is added carefully, the mixture is stirred further for 20 minutes and NH<sub>4</sub>Cl (70g) is added. The cooling bath is removed, the reaction temperature is made to rise to 0°, H<sub>2</sub>O (400ml) is carefully added and the liquid phases are allowed to separate. The aqueous phase is extracted with Et<sub>2</sub>O (3 x 300ml), the combined organic phases are washed with saturated aqueous NaCl (brine), dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated down (to 200ml), and while being kept at 0°, 2 N aqueous HCl (300ml) is added slowly, then the two phases are stirred vigorously for 1 hour at ~25°, H<sub>2</sub>O (500ml) is added and the phases are separated. The organic layer is treated with H<sub>2</sub>O (150ml) and extracted with CH<sub>2</sub>Cl<sub>2</sub> (to remove any cyclohexanone). All the aqueous acidic phases are combined, and carefully (cool if necessary) basified with 2N aqueous NaOH (300ml) and extracted with CH<sub>2</sub>Cl<sub>2</sub> (4 x 300ml); the combined extracts are washed with brine, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated *in vacuo* to give crude ±STEIN (~40g, ~90%) as a light yellow solid *m* 81-82°. Its <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>, TMS) has δ<sub>H</sub> at 1.59 (br s, 4H), 4.10 (s, 2H), 7.2-7.3 (m, 10H); and its <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has δ at 61.9, 126.8, 126.9, 128.2, 143.4.

The optical resolution is carried out by adding carefully a hot (70°) homogeneous solution of (L)-(+)-tartaric acid (30.0g, 0.20mol) in EtOH (230ml) to a hot (70°) solution of ±STEIN (42.5g, 0.20mol) in EtOH (230ml)

(care exothermic, EtOH may boil) with stirring. The tartrate salts separate immediately and the mixture is cooled to  $\sim 25^\circ$ , the salts are filtered off, washed with EtOH (2 x 60ml), dried *in vacuo*, dissolved in boiling H<sub>2</sub>O (230ml), EtOH (230ml) is added and the clear solution is cooled slowly to  $\sim 25^\circ$ . The crystals are collected by filtration, washed with EtOH (40ml), dried *in vacuo*, and recrystallised twice as before [from boiling H<sub>2</sub>O (230ml), and EtOH (230ml)] to give the colourless (-)-*diamine*-(+)-*tartrate salt* ( $\sim 24\text{g}$ ,  $\sim 66\%$ ) with  $[\alpha]_{\text{D}}^{23} -10.8^\circ$  (c 1.2, H<sub>2</sub>O). To this (-)-(+)-salt suspended in H<sub>2</sub>O (300ml) and vigorously stirred at 0-5 $^\circ$ , is added dropwise an aqueous solution of 50% NaOH (23ml) followed by CH<sub>2</sub>Cl<sub>2</sub> (150ml), and is stirred for 30 minutes; the phases are separated, the aqueous phase is extracted with CH<sub>2</sub>Cl<sub>2</sub> (2 x 50ml), the organic layers are combined, washed with brine, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated *in vacuo* to give a colourless solid that is recrystallised from hexane to provide  $>98\%$  optically pure *S,S*-(-)-*diphenylethylenediamine* as colourless crystals with the properties stated in the title and with NMR spectra same as the racemate.

To obtain the enantiomeric diamine, all the filtrates from above are combined, evaporated to dryness *in vacuo*, the residue is stirred vigorously in H<sub>2</sub>O (250ml), treated with aqueous 50% NaOH (25ml) then CH<sub>2</sub>Cl<sub>2</sub> (200ml), stirred further for 30 minutes, the phases are separated, the aqueous phase is extracted with CH<sub>2</sub>Cl<sub>2</sub> (2 x 50ml), the combined extracts are washed with brine, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated *in vacuo* to give enriched *R,R*-(+)-*diamine* ( $\sim 26\text{g}$ ) as pale yellow crystals. This is treated with (D)-(-)-tartaric acid as in the above and gives relatively impure (+)-*diamine*-(-)-*tartrate salt* ( $\sim 31\text{g}$ ,  $\sim 85\%$ ,  $[\alpha]_{\text{D}}^{23} +4^\circ$  (c 1.2, H<sub>2</sub>O). However, recrystallisation of this salt did not improve its optical purity, but treatment with 50% NaOH and extraction as above etc, and crystallisation from hexane gave optically pure *R,R*-(+)-*diphenylethylenediamine* as colourless crystals with  $[\alpha]_{\text{D}}^{23} +106^\circ$  (c 1.1, MeOH), and NMR spectra as obtained with the racemate. The optical purity can be confirmed by the <sup>1</sup>H NMR of their (L)-mandelate salts (*cf* Benson et al. *J Org Chem* **53** 5335 1988). These diamines were used successfully in highly stereoselective epoxidation [Zhang et al. *J Am Chem Soc* **112** 2801 1990], aldol, Diels-Alder [Corey et al. *J Am Chem Soc* **111** 5493 1989], allylation [Corey et al. *J Am Chem Soc* **111** 5495 1989], osmylation [Corey et al. *J Am Chem Soc* **111** 2943 1989] reactions; as well as enantioselective Michael additions [Bruner & Hammer *Angew Chem Inter Ed Eng* **23** *J Am Chem Soc* **111** 5493 1989] 312 1984] and asymmetric hydrogenation [Fiorini & Giongo *J Mol Cat* **5** 303 1979]. [Corey et al. *J Am Chem Soc* **111** 5493 1989, Pikul & Corey *Org Synth* **71** 22 1993 and references herein.]

This diamine has been used successfully as a Trost ligand with Pd(dba)<sub>2</sub> to catalyse a Heck reaction, but it is not, albeit, as commonly used as DACH (*trans*-1,2-diaminocyclohexane) ligands.

***R*-(+) and *S*-(-)-2-[2-(Diphenylphosphino)phenyl]-4-isopropyl-2-oxazoline** [*R*-(+)- and *S*-(-)-2-[2-(diphenylphosphino)phenyl]-4,5-dihydro-4-isopropylloxazole [*R*- 164858-78-0; *S*- 148461-14-7] M 373.4, m 85-89 $^\circ$ ,  $[\alpha]_{\text{D}}^{20}$  *R* +48 $^\circ$  and *S* -48 $^\circ$  (c 1.4, CHCl<sub>3</sub>). The 4*S*-(-) enantiomer can be obtained by decomposing the 4*S*-(+)-Zinc II dichloride complex [crystallises from *tert*-butyl-methyl-ether/CHCl<sub>3</sub> as colourless crystals m 221-223 $^\circ$ ,  $[\alpha]_{\text{D}}^{20} +53^\circ$  (c 1.52, CHCl<sub>3</sub>)] (9.5g, 17.5mmol) in CHCl<sub>3</sub> (130ml) under argon with 2,2'-dipyridyl (2.72g, 1.74mmol) at  $\sim 25^\circ$  for 1 hour (stirring), then the mixture is applied directly onto a silica-gel column (6 x 7 cm), eluted with CHCl<sub>3</sub> or EtOAc (800ml), and the eluate is evaporated to yield the desired 4*S*-(-)-oxazoline (6.78g 95%) as a colourless solid which can be recrystallised from petroleum ether/Et<sub>2</sub>O. On TLC (silica gel 60 Merck, 0.25mm F245) it has R<sub>F</sub> 0.40 (hexane/EtOAc 6/1); the IR (CHCl<sub>3</sub>) has  $\nu_{\text{max}}$  at 3070w, 3060m, 3010m, 2965s, 2905m, 2875m, 1950w, 1885w, 1815w, 1650s, 1585w, 1480m, 1470w, 1435s, 1355m, 1310m, 1245m, 1090s, 1050m, 1030m, 965m cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 0.70 (d, *J* = 6.8Hz, 3H, CH(CH<sub>3</sub>)<sub>2</sub>), 0.81 (d, *J* = 6.8Hz, 3H, CH(CH<sub>3</sub>)<sub>2</sub>), 1.43-1.54 (m, 1H, CH(CH<sub>3</sub>)<sub>2</sub>), 3.80-3.90 (m, 2H, H<sub>2</sub>C(5)), 4.09-4.19 (m, 1H, HC(4)), 6.85-6.89 (m, 1H, HC(4')), 7.24-7.37 (m, 12 H, arom. H), 7.88-7.92 (m, 1H, HC(6')); the <sup>13</sup>C NMR (75MHz, CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 18.3 (CH(CH<sub>3</sub>)<sub>2</sub>), 18.9 (CH(CH<sub>3</sub>)<sub>2</sub>), 32.7 (CH(CH<sub>3</sub>)<sub>2</sub>), 70.0 (H<sub>2</sub>C(5)), 73.1 (HC(4)), 127.8, 128.05, 128.16, 128.26, 128.27 (d, *J*<sub>PC</sub> = 15Hz), 129.6 (d, *J*<sub>PC</sub> = 3Hz), 130.1, 133.5 (d, *J*<sub>PC</sub> = 20Hz), 133.7, 134.1 (d, *J*<sub>PC</sub> = 21Hz)(arom CH), 131.7 (d, *J*<sub>PC</sub> = 19Hz), 137.9 (d, *J*<sub>PC</sub> = 10Hz), 138.1 (d, *J*<sub>PC</sub> = 12Hz), 138.8 (d, *J*<sub>PC</sub> = 25Hz)(arom. H), 162.7 (d, *J*<sub>PC</sub> = 3Hz, C(2)); and the <sup>31</sup>P NMR (121MHz, CDCl<sub>3</sub>, triphenyl phosphate external standard at -18ppm) has  $\delta$  at -5.8.

***S*-(+)-2-[2-(diphenylphosphino)phenyl]-4,5-dihydro-4-phenylloxazole** [148461-15-6] M 407.4, m 89-93 $^\circ$ ,  $[\alpha]_{\text{D}}^{20} +29.5^\circ$  (c 1, CHCl<sub>3</sub>) and ***R*-(-)-2-[2-(diphenylphosphino)phenyl]-4,5-dihydro-4-phenylloxazole** [167171-03-1] M 407.4, m 89-93 $^\circ$ ,  $[\alpha]_{\text{D}}^{20} -29.5^\circ$  (c 1, CHCl<sub>3</sub>) were prepared as for the 4-isopropylloxazole above and have R<sub>F</sub> 0.27 (hexane/EtOAc 5/1). [Koch et al. *Rec Trav Chim Pays Bas* **114** 206 1995.]

With an iridium catalyst they promote asymmetric alkylation [Janssen & Helchen *Tetrahedron Lett* **38** 8025

1997], and are chiral ligands for asymmetric reduction reactions [lightfoot et al. *Angew Chem Int Ed Engl* **110** 3047 1998]. These ligands, when complexed with Ni(COD)<sub>2</sub>, promote the regio and stereoselective addition of allylic sulfides to alkynes; a reaction which tolerates a variety of functional groups [Hua et al. *Org Lett* **9** 263 2007]. These types of ligands also promote Pd-catalysed allylic substitution reactions [von Matt & Pflatz *Angew Chem* **105** 614 1993, Williams et al. *Tetrahedron Lett* **34** 3149 1993].

**Diphenylphosphine** [829-85-6] **M 186.2, b 100-102°/1.5mm, 150-151°/12mm, 156-157°/16mm, 165°/25mm, 280°/atm, d<sup>25</sup> 1.07, n<sub>D</sub><sup>20</sup> 1.6269, pK 0.03.** This toxic and flammable phosphine is prepared under N<sub>2</sub> by adding dropwise a solution of chlorodiphenylphosphine (50g, [1079-66-9]) in dry Et<sub>2</sub>O (75ml) to a solution/suspension of LiAlH (2.6g) in absolute Et<sub>2</sub>O (75ml), then boiling the mixture under reflux for 1 hour, adding H<sub>2</sub>O (6ml, carefully and with cooling) and stirring for 2-3 hours. The Et<sub>2</sub>O layer is collected, dried (CaCl<sub>2</sub>), decanted, evaporated and the residue is distilled *in vacuo* to give the phosphine (70%) as a colourless oil [Kucken & Buchwald *Chem Ber* **91** 2871 1958]. Alternatively, Ph<sub>2</sub>ClP is reacted with Na in dry *n*-Bu<sub>2</sub>O under N<sub>2</sub> to give Ph<sub>2</sub>PNa which is treated with EtOH, boiled under reflux for 30 minutes, poured into H<sub>2</sub>O, extracted with Et<sub>2</sub>O which is separated, dried (CaCl<sub>2</sub>), decanted, evaporated and the residue is distilled *in vacuo* to give the phosphine (72%) [Kucken & Buchwald *Chem Ber* **92** 227 1959]. [Beilstein **16** H 758, **16** II 371, **16** III 833, **16** IV 950.]

**2-(Diphenylphosphino)benzoic acid [(2-carboxyphenyl)diphenylphosphine]** [17261-28-8] **M 306.3, m 174-177°, 175-176° (187-188°), pK<sub>Est</sub> ~2.7.** The benzoic acid is prepared by first adding Na (15.3g, 667mmol, cut in ~1g pieces) carefully to liquid NH<sub>3</sub> (1.5L) with stirring until a blue colour is persistent (**use an efficient fume cupboard**), then adding PPh<sub>3</sub> (87.4g, 330mmol) in small portions during 30-40 minutes with swirling to dissolve any Na on the sides of the flask. After 2.5 hours the red-orange colour of the solution of NaPPh<sub>2</sub> is treated with 2-chlorobenzoic acid (52.2g, 330mmol) in small portions over 30-40 minutes followed by anhydrous THF (500ml, using a syringe), and the resulting solution is allowed to warm to room temperature overnight under a blanket of argon. As the solution warms, the colour of the red solid turns to gold with mild evolution of heat results. The residue is dissolved in H<sub>2</sub>O (1.4L) and extracted with Et<sub>2</sub>O (400ml) which is discarded. The aqueous layer is acidified to pH 2 with concentrated HCl (~60ml), and extracted with CH<sub>2</sub>Cl<sub>2</sub> (3 x 200ml), the extract is washed with H<sub>2</sub>O (500ml), evaporated to *ca* 125ml, and MeOH (~50ml) is added to precipitate the pale-yellow crystalline acid (requiring inducement sometimes) which is collected and dried to give the phosphino-acid (50g, 49%) which provides analytically pure acid (C, H and P) on recrystallisation from a minimum amount of boiling MeOH. The air stable compound forms yellow needles from aqueous EtOH, is soluble in halogenated solvents and Et<sub>2</sub>O, but is less soluble in other solvents. Its IR (Nujol) has  $\nu_{\max}$  at 3200 (OH), 1690 (CO), 1275 cm<sup>-1</sup>. The *methiodide* crystallises from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub> with **m 169-173°(dec)**, and the *phosphine oxide* [2572-40-9] **M 322.3 has m 262-264° (252-254° and 274-275° are also reported), pKa 5.06 and 5.37** (in 50% aqueous EtOH) and crystallises from EtOH, aqueous EtOH or aqueous AcOH. With diazomethane, *methyl 2-(diphenylphosphino)benzoate* is obtained in 85% yield; it forms colourless crystals from MeOH with **m 96-97°**, its IR (Nujol) has  $\nu_{\max}$  at 1712 (C=O) cm<sup>-1</sup>, and the <sup>1</sup>H NMR (90MHz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 7.6-6.9 (m, 14H), 3.70 (s, 3H), and the <sup>31</sup>P NMR (40.5MHz, CDCl<sub>3</sub>) has  $\delta$  at 5.10 upfield of 85% H<sub>3</sub>PO<sub>4</sub>. [Rauchfuss et al. *J Am Chem Soc* **103** 6769 1981, Hoots et al. *Inorg Chem* **21** 178 1982.] It is used for preparing the DACH-Trost phenyl ligand above.

**S(-)-2-Ferrocenyl-4-(1-methylethyl)oxazoline [S(-)-4-isopropylxazo-2-yl ferrocene]** [162157-03-1] **M 297.0, m 71.5-72.5° (m 61-62°), [α]<sub>D</sub><sup>22</sup> -129° (c 1.5, EtOH), CD (CHCl<sub>3</sub>) λ<sub>max</sub> (Δε) 464 (-080), 358 (+1.0)nm.** This complex is prepared in three steps from ferrocenecarboxylic acid (1.33g, 4.49mmol, 1271-42-7), **firstly** by converting it into the acid chloride with oxalyl chloride (0.79g, 9mmol) in CH<sub>2</sub>Cl<sub>2</sub> (15ml) under N<sub>2</sub> at ~25° until gas evolution ceased (~10 minutes), stirring for 10 minutes further, and the solution is evaporated to dryness *in vacuo*. **Secondly** the crude dark red oily acid chloride (that solidified on standing) is dissolved in CH<sub>2</sub>Cl<sub>2</sub> (10ml) and added (syringe) to a solution of S(+)-valinol (0.554g, 5.37mmol) and Et<sub>3</sub>N (1.25ml, 9mmol) in CH<sub>2</sub>Cl<sub>2</sub> (7ml) under N<sub>2</sub> at ~25°. After standing overnight, the mixture is washed with H<sub>2</sub>O (2 x 20ml), dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, and evaporated to dryness *in vacuo*. The residue is purified by column chromatography on SiO<sub>2</sub> (40-83mm), eluting with 3% MeOH/CH<sub>2</sub>Cl<sub>2</sub>, to give analytically pure **2S-N-(1-hydroxy-3-methylbutyl)-ferrocenamido** (1.184g, 84%) with **m 109-110°, [α]<sub>D</sub><sup>21</sup> -8° (c 1.34, EtOH)**. It has IR (nujol) with  $\nu_{\max}$  at 3284 (N-H), 3192 (O-H), 1611 (C=O amide I), 1551 (amide II) cm<sup>-1</sup>; the <sup>1</sup>H NMR (360MHz, CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 0.97

(d, 3H,  $J = 6.8\text{Hz}$ ,  $-\text{CH}_3$ ), 0.98 (d, 3H,  $J = 6.8\text{Hz}$ ,  $-\text{CH}_3$ ), 1.93 (octet, 1H,  $J = 6.8\text{Hz}$ ,  $-\text{CH}(\text{CH}_3)_2$ ), 2.71 (brt, 1H,  $-\text{OH}$ ), 3.56-3.83 (m, 3H,  $-\text{CH}_2\text{OH}$  and  $-\text{NHCH}-$ ), 4.16 (s, 5H,  $\text{C}_5\text{H}_5$ ), 4.30 (brs, 2H, Fc), 4.60 (brs, 1H, Fc), 4.62 (brs, 1H, Fc), 5.80 (brd, 1H,  $-\text{NH}-$ ); the  $^{13}\text{C}$  NMR (90MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}\{^1\text{H}\}$  at 19.04 ( $-\text{CH}_3$ ), 19.69 ( $-\text{CH}_3$ ), 28.99 ( $-\text{CH}(\text{CH}_3)_2$ ), 57.01 ( $-\text{NHCH}-$ ), 63.87 ( $-\text{OCH}_2-$ ), 67.91 (Fc), 68.42 (Fc), 69.75 ( $\text{C}_5\text{H}_5$ ), 70.46 (Fc), 70.51 (Fc), 75.95 (Fc  $-\text{ipso}$ ), 171.37 (C=O); and the MS (EI) has  $m/z$  315 ( $\text{M}^+$ , 100%), 213 (81%). **Thirdly**, to a light orange solution of the preceding *S*-(-)-amide (0.817g, 2.59mmol) and  $\text{Ph}_3\text{P}$  (2.49g, 9.5mmol) in dry MeCN (60ml) is added  $\text{Et}_3\text{N}$  (1.6ml, 11.5mmol) followed by  $\text{CCl}_4$  (2.2ml, 22.8mmol), and the mixture is stirred at  $\sim 25^\circ$  overnight. Water (80ml) is added to the mixture which is extracted with petroleum ether (5 x 50ml), the combined organic layers are dried ( $\text{MgSO}_4$ ), filtered, and evaporated to dryness *in vacuo*. The residue is contaminated with  $\text{Ph}_3\text{PO}$ , and is purified through column chromatography ( $\text{SiO}_2$ , 40-83mm) and eluting with 30% EtOAc/petroleum ether (b 40-60 $^\circ$ ) to give analytically pure ferrocenyl oxazoline as dark yellow crystals (0.685g, 89%) with **m 71.5-72.5 $^\circ$** . Note that much lower yields of product are obtained if the solvent is removed from the reaction mixture before work-up. It has IR (nujol) with  $\nu_{\text{max}}$  at 1657 (C=N)  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (360MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 0.87 (d, 3H,  $J = 6.8\text{Hz}$ ,  $-\text{CH}_3$ ), 0.94 (d, 3H,  $J = 6.8\text{Hz}$ ,  $-\text{CH}_3$ ), 1.78 (hextet, 1H,  $J = 6.6\text{Hz}$ ,  $-\text{CH}(\text{CH}_3)_2$ ), 3.41 (q, 1H,  $J = 7.0\text{Hz}$ ,  $-\text{NHCH}-$ ), 3.89-3.95 (m, 1H,  $-\text{OCHH}-$ ), 4.00 (t, 1H,  $J = 7.7\text{Hz}$ ,  $-\text{OCHH}-$ ), 4.06 (s, 5H,  $\text{C}_5\text{H}_5$ ), 4.26 (brs, 2H, Fc), 4.66 (brs, 1H, Fc), 4.70 (brs, 1H, Fc); the  $^{13}\text{C}$  NMR (90MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}\{^1\text{H}\}$  at 17.79 ( $-\text{CH}_3$ ), 18.81 ( $-\text{CH}_3$ ), 32.27 ( $-\text{CH}(\text{CH}_3)_2$ ), 68.92, 68.95, 69.27, 69.51 ( $\text{C}_5\text{H}_5$ ), 70.06, 70.09, 70.60 (Fc  $-\text{ipso}$ ), 72.27, 165.56 (C=N); and the MS (EI) has  $m/z$  297 ( $\text{M}^+$ , 100%), 254 (64%), 211 (47%), 121 (92%). [Richards & Mulvaney *Tetrahedron Asymm* 7 1419 1996.] The *R*-enantiomer is prepared in the same way by using *R*(+)-valinol. Similarly prepared while using the respective chiral *S*-2-aminoethanols were *S*-(-)-4-methyloxazol-2-yl ferrocene [red solid, 96% yield, **m 84-85 $^\circ$** ,  $[\alpha]_{\text{D}}^{27} -60^\circ$  (c 0.2, EtOH), CD ( $\text{CHCl}_3$ )  $\lambda_{\text{max}}$  ( $\Delta\epsilon$ ) 461 (+0.34), 336 (-0.25)nm], *S*-(-)-4-ethyloxazol-2-yl ferrocene [orange solid, 34% yield, **m 47-48 $^\circ$** ], *S*-(-)-4-tert-butyloxazol-2-yl ferrocene [orange-red solid, 34% yield, **m 127-128 $^\circ$** ,  $[\alpha]_{\text{D}}^{24} -150^\circ$  (c 2.1,  $\text{CH}_2\text{Cl}_2$ )], and among others where the 4-alkyl group is replaced by isobutyl, *sec*-butyl, benzyl, phenyl, 4,5-diphenyl, 1-hydroxy-1-methylethyl, 1-methoxy-1-methylethyl, 1-hydroxy-1-ethylpropyl, 1-methoxy-1-ethylpropyl, hydroxymethyl, methoxymethyl, and several more [Sammakia et al. *J Org Chem* 60 10 1995, Nishibayashi et al. *J Organomet Chem* 545-546 381 1997, Richards & Mulvaney *Tetrahedron Asymm* 7 1419 1996.] An independent synthesis of the title compound from *S*-(-)-valinol,  $\text{ZnCl}_2$ , cyanoferrocene in refluxing chlorobenzene under  $\text{N}_2$  for 72 hours followed by column chromatography ( $\text{SiO}_2$ , eluting with hexane/EtOAc) gave a red solid (40% yield based on cyanoferrocene) **m 61-62 $^\circ$**  [note difference in m.p.] but with similar  $^1\text{H}$  and  $^{13}\text{C}$  NMR. Other chiral derivatives, some mentioned above, were also obtained in this way. [Nishibayashi et al. *J Organomet Chem* 545-546 381 1997.]

**Iridium(III) acetylacetonate [tris(pentane-2,4-dionato)iridium,  $\text{Ir}(\text{acac})_3$ ] [15635-87-7] M 489.5, m 269-271 $^\circ$** . Iridium sulfate solution is made from freshly precipitated iridium dioxide (from 1.0g of potassium hexachloridate) by heating it with 1N  $\text{H}_2\text{SO}_4$  (25ml) until the acid begins to fume. The solution is diluted to its original volume, and undissolved iridium oxide is removed by centrifugation. The light green solution is treated with 10% aqueous NaOH until the green iridium III hydroxide precipitate just begins to redissolve. This precipitation should be rapid in order to avoid oxidation; the pH of the solution is adjusted to 6, and the mixture is heated at 60 $^\circ$  with acetylacetone (2ml) for one hour. The solution becomes red and deposits a yellow crystalline precipitate. This iridium complex is collected and crystallises from aqueous MeOH in rhombic plates (0.1g, 10%, m 269 $^\circ$ ). It is insoluble in  $\text{H}_2\text{O}$ , slightly soluble in EtOH and petroleum ether, but freely soluble in  $^*\text{C}_6\text{H}_6$  and  $\text{CHCl}_3$ . It is stable in boiling dilute acids and 10% aqueous NaOH. It sublimes at 260 $^\circ$ /1mm, but decomposes above 290 $^\circ$  depositing an iridium mirror. Molecular weight determination ( $\sim 400$ , by Rast in camphor) indicates that it is monomeric. [Dwyer & Sargeson *J Am Chem Soc* 75 984 1953, Beilstein 1 IV 3678.] It is a useful precursor for making Iridium complexes.

**3,4-O-isopropylidene-(3*S*,4*S*)-dihydroxy-(2*R*,5*R*)-bis(diphenylphosphino)hexane (RSSR-dimeDIOP, or *R,S,S,R*-DIOP\*, DIOP starred, in dedication to Professor Kagan) [258873-45-9] M 526.2, low melting oil,  $[\alpha]_{\text{D}}^{24}$  (+) 41.8 $^\circ$  (c 0.88,  $\text{PhCH}_3$ ),  $\text{pK}_{\text{Est}} \sim 0.0$ . The starting material for this ligand is the readily available D-mannitol that is converted to the isopropylidene diepoxide, which in turn is converted to 3,4-*O*-isopropylidene-(3*S*,4*S*)-dihydroxy-(2*S*,5*S*)-hexane. The latter (2.2g, 11.6mmol), and  $\text{Et}_2\text{NH}$  (4.9ml, 34.8mmol) in  $\text{CH}_2\text{Cl}_2$  (30ml) is added dropwise to a solution of  $\text{MeSO}_2\text{Cl}$  (2.0ml, 25.8mmol) in  $\text{CH}_2\text{Cl}_2$  (10ml) at 0 $^\circ$ , kept at 0 $^\circ$  for 30 minutes, stirred at  $\sim 25^\circ$  for 30 minutes, and treated with saturated aqueous  $\text{NH}_4\text{Cl}$ . The layers are separated, the**

aqueous phase is extracted with  $\text{CH}_2\text{Cl}_2$ , the combined organic phases are dried ( $\text{Na}_2\text{SO}_4$ ), evaporated and the residue is purified by flash chromatography on silica gel, eluting with  $\text{CH}_2\text{Cl}_2/\text{Et}_2\text{O}$  (9/1) to give **3,4-O-isopropylidene-(3S,4S)-bis(mesyloxy)-(2S,5S)-hexane** (3.85g, 96%) as a colourless oil. This is a useful intermediate for making a variety of phosphine ligands and has  $[\alpha]_{\text{D}}^{24}$  (-) **1.3°** (c **1.04**,  $\text{CHCl}_3$ ); its  $^1\text{H}$  NMR (360MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 4.82-4.76 (m, 2H), 3.99-3.96 (m, 2H), 3.03 (s, 6H), 1.45 (d,  $J = 6.6, 6.0\text{Hz}$ ), 1.37 (s, 6H); and the  $^{13}\text{C}$  NMR (90MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 110.14, 78.19, 76.26, 38.53, 26.75, 17.63; and HRMS has found  $m/z$  347.0834, calcd for  $\text{C}_{11}\text{H}_{23}\text{O}_8\text{S}_2$  ( $\text{MH}^+$ ) is 347.0834. The preceding dimesylate (1.04g, 3.0mmol) in THF (20ml) is added to the orange solution at  $-78^\circ$  of  $\text{Ph}_2\text{PLi}$  [prepared under  $\text{N}_2$ , from  $\text{Ph}_2\text{PH}$  (1.15ml, 6.6mmol) in THF (50ml) to which is added *via* a syringe *n*-BuLi in hexane (4.0ml, 6.4mmol) at  $-78^\circ$  during 5 minutes, warmed to  $\sim 25^\circ$  while stirring for 1 hour, and cooling to  $-78^\circ$  again] during 20 minutes. The orange solution is warmed to  $\sim 25^\circ$ , stirred overnight, the resulting white suspension is hydrolysed with saturated aqueous  $\text{NH}_4\text{Cl}$ , the aqueous layer is extracted with  $\text{CH}_2\text{Cl}_2$ , the combined organic layers are dried ( $\text{Na}_2\text{SO}_2$ ), filtered, evaporated and the residue is purified by flash chromatography on silica gel, eluting with hexane/ $\text{Et}_2\text{OAc}$  (95/5), or hexane/ $\text{Et}_2\text{O}$  (95/5) to give (+)-RSSR-DIOP\* (1.06g, 67%) as a colourless oil or low melting white solid respectively. Its  $^1\text{H}$  NMR (360MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 7.56-7.52 (m, 8H), 7.38-7.33 (m, 12H), 3.78-3.76 (m, 2H), 2.50-2.46 (m, 2H), 1.14 (s, 6H), 0.91 (d,  $J = 7.0\text{Hz}$ , 3H), 0.87 (d,  $J = 6.9\text{Hz}$ , 3H); the  $^{13}\text{C}$  NMR (90MHz,  $\text{CD}_2\text{Cl}_2$ ) has  $\delta$  at 137.46 (d,  $J = 15.9\text{Hz}$ ), 137.03 (d,  $J = 15.5\text{Hz}$ ), 134.14 (d,  $J = 3.7\text{Hz}$ ), 133.91 (d,  $J = 4.0\text{Hz}$ ), 129.30 (d,  $J = 8.6\text{Hz}$ ), 128.75 (d,  $J = 7.1\text{Hz}$ ), 108.27 (dd,  $J = 12.0, 6.8\text{Hz}$ ), 31.33 (d,  $J = 14.3\text{Hz}$ ), 27.05, 10.74 (d,  $J = 17.6\text{Hz}$ ); the  $^{31}\text{P}$  NMR ( $\text{CDCl}_3$ ) has  $\delta$  at -6.3 (-5.42); and the HRMS has found  $m/z$  527.2771, calculated for  $\text{C}_{33}\text{H}_{37}\text{O}_2\text{P}_2$  ( $\text{MH}^+$ ) is 527.2269. [Li & Zhang *J Org Chem* **65** 5871 2000, Yan & RajanBabu *Org Lett* **26** 4137 2000.]

**Lithium tetrachloropalladate(II) hydrate** ( $\text{Li}_2\text{PdCl}_4 \cdot x\text{H}_2\text{O}$ ) [123334-21-4] **M 262.1 (anhydrous)**. It is commercially available as a powder or in chunks. A 0.5M solution of water-free  $\text{Li}_2\text{PdCl}_4$  in dry MeOH for the preparation of Pd catalysts (e.g. see *Oxime Palladacycles* above) is made by stirring anhydrous LiCl (42.4g, 1 mole, dried at  $130^\circ$  to constant weight, several hours, see [7447-41-8]) and anhydrous  $\text{PdCl}_2$  (88.6g, 0.5 mole, see [7647-10-1]) in dry MeOH (1L) until the solution is homogeneous. Store the solution under  $\text{N}_2$  or argon. [Onue, Minami and Nakagawa *Bull Chem Soc Jpn* **43** 3480 1970.]

**$\eta^5$ -(S)-(pR)-2-(2'-(4'-Methylethyl)-oxazoliny)cyclopentadienyl,1-C-3'-N)( $\eta^4$ -tetraphenylcyclobutadiene) cobalt** [222400-02-4] **M 591.6, m 160-162°**,  $[\alpha]_{\text{D}}^{24}$  -55.2° (c **0.09**,  $\text{CHCl}_3$ ). This cobalt complex is the key intermediate for the preparation of all the cobalt oxazoline palladacycle (COP) catalysts described in Part 1. Although originally prepared in four distinct steps [Stevens & Richards *Organometallics* **18** 1346 1999], the synthesis has been slightly modified and can be completed in two steps without purifying intermediates. **Firstly**,  **$\eta^5$ -(carbomethoxycyclopentadienyl)-( $\eta^4$ -tetraphenylcyclobutadiene)cobalt** [182627-81-2] is prepared (using Schlenk equipment flushed with argon) in a triple-necked flask with two rubber septa containing sodium cyclopentadienide (20ml, 2.0M in THF, 40mmol, [4984-82-1], should be light red/orange in colour; discard if black or contains solid residues), anhydrous THF (20ml *via* a syringe, pre-filtered through activated  $\text{Al}_2\text{O}_3$ ), then dimethyl carbonate (10.2ml, 120mmol, *via* a syringe), and the mixture is stirred and boiled under reflux for 4 hours before cooling to  $\sim 25^\circ$ . [The intermediate can be isolated but this is not necessary.] To the cooled mixture is added dry toluene (160ml, *via* a syringe), followed by chlorotris(triphenylphosphine)cobalt(I) (30.8g, 34.8mmol, see above [26305-75-9], preferably freshly prepared for higher yields) and diphenylacetylene (14.2g, 80mmol, recrystallised from EtOH), and heated at reflux for 5 hours, cooled to  $\sim 25^\circ$  before transferring to a single-necked flask and evaporating to dryness *in vacuo*. The residue is suspended in hexanes (150ml), the solid is filtered off, washed with hexanes ( $\sim 3.5\text{L}$ ) until the filtrate is colourless. The dark mustard-coloured solid is dissolved from the filter by washing with  $\text{CH}_2\text{Cl}_2$  until the filtrate is colourless (discard any insoluble black solid) and evaporate the combined filtrates *in vacuo* to give the **Cobalt-ester** (12.7g, 67%) as a mustard-coloured solid. It can be purified by column chromatography on Matrix silica 60 (35-70 $\mu\text{m}$ , eluting with 20% EtOAc/petroleum ether b 40-60°) to give the **Cobalt-ester** as an orange crystalline solid **m 240-244°** [lit **m 246-248°** Mabrouk & Rauch *J Organomet Chem* **523** 111 1996]. It has IR (thin film) with  $\nu_{\text{max}}$  at 3058, 2953, 1713, 1596, 1498, 1467, 1281, 1140  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (500MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 3.23 (s, 3H), 4.79 (t,  $J = 2.0\text{Hz}$ , 2H), 5.21 (t,  $J = 2.0\text{Hz}$ , 2H), 7.22-7.30 (m, 12H), 7.43-7.45 (m, 8H); and the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 51.4, 76.5, 84.7, 86.6, 86.8, 126.9, 128.2, 129.0, 135.3, 166.6. [Anderson, Overman, Richards, Watson and

White *Org Synth* **84** 139 2007.]

**Secondly**,  $\eta^5$ -(*S*)-(p*R*)-2-(2'-(4'-methyleneethyl)oxazoliny)cyclopentadienyl, 1-*C*-3'-*N*)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt is prepared by stirring under argon the above crude *Cobalt-ester* (10g, 18.6mmol), lithium iodide (4.96g, 37.2mmol, see [10377-51-2]) and 2,4,6-collidine (100ml, b **175-178**<sup>o</sup>/atm, see [108-75-8]), then boiling under reflux (air condenser) for ~16 hours until the *Cobalt-ester* ( $R_F$  0.8, silica plate eluted with 20% EtOAc:hexanes) has completely reacted. The solution is cooled to ~25<sup>o</sup>, diluted with CH<sub>2</sub>Cl<sub>2</sub>, washed with 2N aqueous HCl (4 x 150ml), dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated *in vacuo* to give the solid orange coloured *Cobalt-acid* [<sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) at  $\delta_H$ : 4.84 (2d,  $J = 2.0$ Hz, 2H), 5.23 (2d,  $J = 2.0$ Hz, 2H), 7.20-7.24 (m, 12H), 7.42-7.44 (m, 8H)] which is used further. Under argon, with stirring, the preceding *cobalt-acid* in CH<sub>2</sub>Cl<sub>2</sub> (124ml, pre-filtered through activated Al<sub>2</sub>O<sub>3</sub>), oxalyl chloride (3.25ml, 37.2mmol) followed by anhydrous DMF (3 drops sequentially, as evolution of gas occurs) are kept at ~25<sup>o</sup> for 30 minutes, evaporated *in vacuo*, CH<sub>2</sub>Cl<sub>2</sub> (100ml) is added to dissolve the residue, evaporate again *in vacuo* and the process is repeated three times to give dry *Cobalt-acid chloride* as a red-brown residue. A Schlenk flask with a rubber septum, flushed with argon *via* inlet and outlet needles, is loaded with *S*-valinol hydrochloride (3.6g, 26mmol, prepared from the free base, see [2026-48-4], by treating with 2N HCl in Et<sub>2</sub>O), then dry Et<sub>3</sub>N (15.5ml, 112mmol) and CH<sub>2</sub>Cl<sub>2</sub> (86ml) are added *via* the septum, followed, *via* a cannula, a solution of the preceding crude *Cobalt-acid chloride* in CH<sub>2</sub>Cl<sub>2</sub> (100ml). The mixture is stirred at ~25<sup>o</sup> for 2 hours and cooled to 0<sup>o</sup> in an ice-bath. MeSO<sub>2</sub>Cl (3.6ml, 47mmol, [124-63-0]) is then added to the mixture in one portion *via* a syringe and allowed to warm to ~25<sup>o</sup>. After 16 hours the mixture is washed with saturated aqueous NaHCO<sub>3</sub> (150ml), brine (150ml), the organic layer is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, and concentrated *in vacuo*. The residue is purified by dissolving it in the minimum volume of CH<sub>2</sub>Cl<sub>2</sub> (~50ml), adding it to a silica gel column and eluting with hexanes/EtOAc (9:1) to give  $\eta^5$ -(*S*)-(p*R*)-2-(2'-(4'-methyleneethyl)-oxazoliny)cyclopentadienyl, 1-*C*-3'-*N*)( $\eta^4$ -tetraphenylcyclobutadiene)cobalt (8.2g, 75%) as a yellow solid, which can be recrystallised to analytical purity from 10% EtOAc/petroleum ether to m 160-162<sup>o</sup>. It has IR (thin film) with  $\nu_{max}$  at 3058, 2958, 1652 (C=N), 1597, 1499, 1373, 1113, 1024 cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) has  $\delta_H$  at 0.77 (d,  $J = 6.7$ Hz, 6H, CH<sub>3</sub>), 0.97 (d,  $J = 6.7$ Hz, 3H, CH<sub>3</sub>), 1.40 (oct,  $J = 6.7$ Hz, 1H, CH(CH<sub>3</sub>)<sub>2</sub>), 3.41-3.56 (m, 3H, -CH-CH<sub>2</sub>-), 4.62 (brs, 1H, CpH), 4.80 (brs, 1H, CpH), 5.09 (brs, 1H, CpH), 5.20 (brs, 1H, CpH), 7.17-7.27 (m, 12H, *m+p*-PhH), 7.44-7.47 (m, 8H, *o*-PhH); and the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta_C$ {<sup>1</sup>H} at 20.4 (CH<sub>3</sub>), 26.1 (CH<sub>3</sub>), 35.0 (CH(CH<sub>3</sub>)<sub>2</sub>), 71.5 (-CH-CH<sub>2</sub>-), 74.7 (-CH-CH<sub>2</sub>-), 78.9 (C<sub>4</sub>Ph<sub>4</sub>), 84.1 (*ipso*-CpC), 86.4 (CpC), 87.0 (CpC), 87.1 CpC), 88.4 CpC), 126.8 (*p*-PhC), 128.3 (PhC), 129.3 (PhC), 135.8 (*ipso*-PhC), 162.6 (C=N); and MS(EI) at  $m/z$  591 (M<sup>+</sup>, 100), 415 (82), 356 (8), 178 (60). The ***R*-enantiomer** is prepared in precisely the same way except that *R*-valinol hydrochloride is used, and it differs only in having optical rotation of opposite sign. [Stevens & Richards *Organometallics* **18** 1346 1999, Anderson, Overman, Richards, Watson and White *Org Synth* **84** 139 2007.]

***N*-Methyliminodiacetic acid (MIDA)** [4408-64-4] **M 147.1, m 220<sup>o</sup> (dec), 223-225<sup>o</sup> (dec), 226<sup>o</sup> (dec), pKa** (thermodynamic) **acidic: 2.138 (0<sup>o</sup>), 2.142 (10<sup>o</sup>), 2.146 (20<sup>o</sup>), 2.150 (30<sup>o</sup>), 2.154 (50<sup>o</sup>); basic: 10.474 (0<sup>o</sup>), 10.287 (10<sup>o</sup>), 10.088 (20<sup>o</sup>), 9.920 (30<sup>o</sup>), 9.730 (40<sup>o</sup>)**. It has been prepared from MeNH<sub>2</sub> and formaldehyde cyanohydrin followed by hydrolysis of the dinitrile formed [Eschweiler *Justus Liebigs Ann Chem* **279** 29 1894] and in 63-71% yield from chloroacetic acid and MeNH<sub>2</sub> [Berchet *Org Synth Coll Vol II* 397 1943]. However, a more convenient preparation is from iminodiacetic acid [142-73-4] (13.3g), formic acid (4.6g), aqueous formaldehyde (10ml, 36%) and H<sub>2</sub>O (5ml) at 100<sup>o</sup> until evolution of CO<sub>2</sub> ceases, and after 2 hours the mixture is refluxed for 10 minutes, diluted with excess EtOH, cooled and the precipitated acid is filtered off. Recrystallisation from aqueous EtOH (charcoal) provides pure colourless *MIDA* (14.7g, 90%). [Childs et al. *J Chem Soc* 2174 1948]. It has been purified *via* the Ba salt [Berchet *Org Synth Coll Vol II* 397 1943] and the stability constants of the complexes with Cu<sup>2+</sup>, Mg<sup>2+</sup>, Ca<sup>2+</sup>, Sr<sup>2+</sup>, Ba<sup>2+</sup>, Zn<sup>2+</sup>, Cd<sup>2+</sup>, Hg<sup>2+</sup>, and Ni<sup>2+</sup> have been determined [Swarzenbach et al. *Helv Chim Acta* **38** 1147 1955]. The **diethyl ester** has **b 92-94<sup>o</sup>/0.04mm** [Viscontini et al. *Helv Chim Acta* **35** 451 1952], and the **monoamide** has **m 168<sup>o</sup>** after crystallisation (needles) from aqueous EtOH [Eschweiler *Justus Liebigs Ann Chem* **279** 29 1894]. *MIDA* (**m 215-216<sup>o</sup> dec**) has been prepared on a ~100g scale by the method described by Childs et al. (above) with slight modification in 88% yield, and had IR (thin film) with  $\nu_{max}$  at 2998, 2955, 1682 (C=N), 1477, 1380, 1328, 1223, 1172, 1126, 1065, 1018, 982, 958, 903, 886, 723 cm<sup>-1</sup>; the <sup>1</sup>H NMR (400MHz, D<sub>2</sub>O) has  $\delta_H$  at 2.98 (s, 3H), 3.96 (s, 4H); the <sup>13</sup>C NMR (100MHz, 95:5 DMSO-*d*<sub>6</sub>: D<sub>2</sub>O, TMS) has  $\delta_C$  at 170.0, 56.7, 41.7; the LRMS (ESI<sup>+</sup>) has  $m/z$  (relative intensity) 219.1 (24%), 148.1 (M<sup>+</sup>, 100%), 102.1 (8%); and the HRMS (ESI<sup>+</sup>) for C<sub>5</sub>H<sub>10</sub>NO<sub>4</sub> [M+H<sup>+</sup>] has calculated 148.0604, and found 148.0603. [Ballmer, Gillis and Burke *Org Synth* **86** 344 2008]. [Beilstein 4 H

367, 4 II 800, 4 III 1177, 4 IV 2428].

**MIDA boronic esters:** MIDA is a very useful protecting group for boronic acids and forms stable (more stable than the free acid) *boronates* which are readily prepared from the boronic acid, MIDA in toluene-DMSO (2.5:1) under reflux using Dean-Stark conditions to eliminate H<sub>2</sub>O requiring ~3 hours. When these boronic acid surrogates are suitably substituted on boron they can be made to undergo Suzuki-Miyaura, Stille, Heck, Sonogashira and Negishi coupling, as well as Miyaura borylations. The boronates are easy to handle, stable in air, stable under cross-coupling conditions, surviving temperatures of up to 80°, and acidic reagents like trifluoromethanesulfonic, acetic acid and oxidizing agents such as Jones reagent. The MIDA group can be easily hydrolysed by mild aqueous base (see below) thus also allowing the use of controlled iterative cross-coupling (ICC) of protected haloboronic acids. [Gillis & Burke *Aldrichimica Acta* **42** 17 2009]. A variety of MIDA boronate synthons are available commercially.

**Hydrolysis of 4-(*p*-tolyl)phenylboronic acid MIDA ester (m 214-216°)** to the boronic acid is given here as a typical example for the hydrolysis procedure. The MIDA ester (10.1g, 31.3mmol, 1 equivalent) in THF (220ml), and aqueous 1M NaOH (93.5ml, 93.5mmol, 2.99 equivalents) gives a biphasic mixture with a clear colourless lower layer and a milky white upper layer. The container is capped, and the mixture is stirred vigorously at ~25° for 10 minutes whereby the lower colourless layer becomes clear and the upper layer becomes clear yellow. A saturated aqueous solution of NH<sub>4</sub>Cl (250ml) is added (to act as acidifier) and the mixture is stirred vigorously for 5 minutes, shaken with Et<sub>2</sub>O (4 x 50ml), and the aqueous layer is extracted with freshly prepared THF/Et<sub>2</sub>O (1:1, 400ml). The combined organic layers are dried (MgSO), filtered through Celite, and concentrated by rotary evaporation in a vacuum (40°/20mm), and residual solvent is removed by three azeotropic cycles with MeCN (3 x 50ml) in a rotary evaporator (40°/20mm), and then at higher vacuum (25°/1mm) for 12 hours to give the boronic acid as a fine off-white powder (6.24g, 94% yield, **m 136-138° dec**) with the expected IR, NMR and Mass spectra which indicate ~92% purity. [Ballmer, Gillis and Burke *Org Synth* **86** 344 2008].

**MIDA-disodium salt** is prepared from MIDA (147g, 1 mol), which is dissolved (caution: exothermic) in NaOH (120g, 3 mol) in H<sub>2</sub>O (300ml) to give a clear light yellow solution that is evaporated *in vacuo*. The residue is treated with MeOH (300ml), heated to reflux, filtered through a glass frit, heated to reflux and filtered again, and the process repeated once more. The filtrate is co-evaporated with MeCN (3 x) and lyophilised to give analytically pure **diNa salt** (173g, 91%) as a white powder with <sup>1</sup>H NMR (300MHz, CD<sub>3</sub>OD, TMS) δ<sub>H</sub> at 2.95 (s, 4H), 2.24 (s, 3H); and <sup>13</sup>C NMR (125MHz, CD<sub>3</sub>OD, TMS) δ<sub>C</sub> at 179.2, 63.7, 43.9. It is used for preparing MIDA-boronate esters [Uno et al. *Tetrahedron* **65** 3130 2009, cf also Coombs & Margerum *Inorg Chem* **9** 1711 1970, Davis & Richardson *Inorg Chem* **23** 184 1984, Stringfield & Shepherd *Inorg Chim Acta* **309** 28 2000].

**Norbornadiene (NBD, bicycle[2.2.1]hepta-2,5-diene) [121-46-0] M 92.1, m -19.1°, b 89.5-90°/atm, 90.3°/atm, d<sub>4</sub><sup>20</sup> 0.9064, n<sub>D</sub><sup>25</sup> 1.4684.** NBD is a versatile diene that coordinates with metals, helps in solubilising them in organic solvents, and is used for making a variety of pre-catalysts (see further in this chapter under “Homogeneous Catalysts” Part 1). It is prepared by a general method whereby *cis*- or *trans*- 1,2-*bis*(phenylsulfonyl)ethylene (0.22g, 0.714mmol; *trans* [963-16-6] M 308.4, m 221-223°, 226.5-227°, 229.5° from EtOH or AcOH; *cis*-isomer has m 89.5-90°, 101-101.5°, from EtOH, *Beilstein* **6** IV 1500, Truce & McManimie *J Am Chem Soc* **75** 1672 1953) and freshly distilled cyclopentadiene (1.6g, 2.42mmol) in toluene (5ml) are heated at 130° for 24 hours, evaporated to dryness *in vacuo*, and the residue is recrystallised from \*C<sub>6</sub>H<sub>6</sub>-EtOH (1:1) to give the Diels-Alder adduct 3,6-*methano*-4,5-*bis*(phenylsulfonyl)cyclohex-1,2-ene, m 257-258°, in good yield. Lower yields are obtained in \*C<sub>6</sub>H<sub>6</sub> (~25°/40 hours). Elimination of the β-disulfone groups is achieved by stirring a mixture of the *bis*(diphenylsulfonyl) compound (2.5mmol) and NaH<sub>2</sub>PO<sub>4</sub>·2 H<sub>2</sub>O (5g) in MeOH (~40ml) at room temperature overnight with freshly prepared sodium amalgam (from *ca* 0.4g, 17mmol of Na, [11110-52-4]). The mixture is extracted with pentane (b 36°/atm) and the extract is fractionated at atmospheric pressure to give NBD in 65% yield. Store and stabilise with 0.50.25% of BHT under N<sub>2</sub>. [De Lucchi & Modena *JCS Chem Commun* 914 1982, cf Parham & Heberling *J Am Chem Soc* **77** 1175 1955.] Industrially it is prepared from cyclopentadiene and acetylene at 340°/6atm [Shell Devel. Co., USP 2 875256 1953].

NBD has UV with λ<sub>max</sub> (logε) at 198.5 (3.78), 202infl (3.41), 213infl (3.24), 219infl (3.01), 227infl (2.41)nm (in EtOH), and 194.5 (3.53)nm (in cyclohexane) [Stich, Rotzler & Reichstein *Helv Chim Acta* **42** 1480 1959]; its IR (thin film) with ν<sub>max</sub> of the more intense bands at 2987, 1543, 1311, 1228, 875, 799, 727, 502 cm<sup>-1</sup>, and in CS<sub>2</sub>

and  $\text{CCl}_4$  it has  $\nu_{\text{max}}$  at 3080(s), 2960(vs), 2880(s), 1646(m), 1452(ms), 1335(ms), 1310(vs), 1271(m), 1229(s), 1206(s), 1150(s), 1105(ms), 1063(m), 1016(m), 935(s), 911(s), 885(sh), 870(vs), 795(sh), 715(vs), 650(s)  $\text{cm}^{-1}$  [Abel, Bennett and Wilkinson *J Chem Soc* 3178 1959]; the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 1.99(s, 2H,  $\text{CH}_2$ ), 3.56(s, 2H, bridgehead H), 6.74(s, 4H, olefin H) from TMS, and the  $^{13}\text{C}$  NMR (15MHz,  $\text{CDCl}_3$ , TMS) has  $\delta_{\text{C}}$  at 143.3, 75.3, 50.2. The photoelectron spectrum shows intense bands at 8.69, 9.55, 11.26, 12.51—12.75, 14.24 and weak bands at 15.66—17.16 eV; and the interaction between the two non-conjugated  $\pi$ -bonds has been shown to be 0.85eV [Bischof et al. *Helv Chim. Acta* 52 1745 1969].

**NBD-silver nitrate** complex is formed by shaking for 5 minutes a mixture of  $\text{AgNO}_3$  (1.36g) in  $\text{H}_2\text{O}$  (10ml) and NBD (3ml). The white complex is collected and recrystallised from EtOH to give pure **NBD-(AgNO<sub>3</sub>)<sub>2</sub>** (0.95g) which decomposes in air, in  $\text{H}_2\text{O}$ , and on heating to give a strong odour of NBD. Store it cold and under  $\text{N}_2$ . It is soluble in warm MeOH, EtOH,  $\text{CCl}_4$ ,  $\text{CHCl}_3$ , and  $^*\text{C}_6\text{H}_6$ , but insoluble in  $\text{Me}_2\text{CO}$ ,  $\text{Et}_2\text{O}$  and light petroleum. Its IR (Nujol and hexachlorobutadiene mulls) has  $\nu_{\text{max}}$  at 3020(w), 2960(w), 2885(w), 1470(m), 1385(v, brs), 1325(s), 1310(s), 1243(m), 1189(m), 1046(m), 1014(w), 969(m), 950(vw), 929(m), 890(w), 877(w), 808(m), 795(w), 790(w), 725(ms)  $\text{cm}^{-1}$  [Abel, Bennett and Wilkinson *J Chem Soc* 3178 1959].

Similarly the **NBD-cuprous bromide** complex is formed by shaking for 5 minutes a mixture of anhydrous CuBr (2.6g) in EtOH (30ml) and NBD (5ml). The white crystalline complex is filtered off, and washed with EtOH (5ml) and  $\text{Et}_2\text{O}$  (5ml) to give **NBD-(CuBr)<sub>2</sub>** as fine white crystals which slowly become green in air. The complex, which is insoluble in most organic solvents, leaves CuBr under reduced pressure, and is decomposed by  $\text{H}_2\text{O}$  to give  $\text{Cu}_2\text{O}$ . Its IR (Nujol and hexachlorobutadiene mulls) has  $\nu_{\text{max}}$  at 3027(vw), 3000(vw), 2930(m), 2860(w), 1471(w), 1453(m), 1310(ms), 1265(w), 1245(vw), 1234(w), 1180—1080(br-w), 1050(vw), 993(m), 976(m), 950(w), 938(w), 918(ms), 889(w), 867(w), 777(w), 765(w), 739(m), 719(m)  $\text{cm}^{-1}$  [Abel, Bennett and Wilkinson *J Chem Soc* 3178 1959]. [*Beilstein* 5 IV 879.]

For the *bis(phenylsulfonyl)ethylenes* see also Truce & McManimie *J Am Chem Soc* 76 5745 1954, Montaniari *Gazetta* 86 429 1956, Parham & Heberling *J Am Chem Soc* 77 1175 1955, Modena & Montaniari *Gazetta* 86 436 1956, Adams & Ferretti *J Am Chem Soc* 81 4931 1959, and for UV see Angeletti & Montaniari *Boll Scient Fac Chim Ind Univ Bologna* 15 44, 45 1957; in EtOH the *cis-isomer* has  $\lambda_{\text{max}}$  at 241–243nm, and the *trans-isomer*  $\lambda_{\text{max}}$  at 246nm.

**Pentaphenylferrocenyl(di-tert-butyl)phosphine** [**P(C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>Ph<sub>5</sub>)(t-butyl)<sub>2</sub>**, **Q-Phos**] [312959-24-3] **M 710.7, m 211-219°**. This catalyst ligand is prepared in two steps in high yield, and all reactions are performed in a drybox. Firstly (*di-tert-butylphosphino*)ferrocene is prepared by adding *t*-BuLi (31.6ml, 53.8mmol) in THF to a solution of  $\text{Cp}_2\text{Fe}$  (10.0g, 53.8mmol) in THF (25ml) during 5 minutes at 0°, stirred for 20 minutes and the solvent is removed *in vacuo*. The residue is dissolved in pentane (100ml) and THF (5ml), (*t*-Bu)<sub>2</sub>PCl (5.33g, 29.5mmol) is added and the mixture is stirred for 3 hours after which time degassed MeOH (1ml) is added, and most of the solvents are removed *in vacuo*. The residue is filtered through a plug of silica gel under  $\text{N}_2$  and unreacted  $\text{Cp}_2\text{Fe}$  is eluted with pentane first and then the phosphine is eluted all at once with  $\text{Et}_2\text{O}$ , the solvent is removed *in vacuo*, and the required *phosphine* is crystallised from pentane (yield 7.58g, 78%). In the second step this *phosphine* (1.00g, 3.03mmol),  $\text{Pd}(\text{OAc})_2$  (0.35g, 0.156mmol) and *t*-BuONa (2.93g, 30.5mmol) are dissolved in PhCl (34.10g, 303.0mmol) and heated at 110° for 18 hours. The mixture is filtered through Celite, the PhCl is removed *in vacuo*, and the residue is subjected to chromatography on Silica gel, and eluting with pentane/ $\text{Et}_2\text{O}$  (80:1), to give **P(C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>Ph<sub>5</sub>)(t-butyl)<sub>2</sub>** (1.47g, 68%) as a pink-red solid. Its  $^1\text{H}$  NMR ( $\text{C}_6\text{D}_6^*$ ) has resonances at  $\delta$ : 1.07(d, 11.0Hz, 18H,  $\text{Me}_3$ ), 4.42(t, 1.7Hz, 2H,  $\text{C}_5\text{H}_4$ ), 4.67(d, 1.0Hz, 2H,  $\text{C}_5\text{H}_4$ ), 6.95–6.97(m, 15H, *m,p*- $\text{C}_6\text{H}_5$ ), 7.44–7.48(m, 10H, *o*- $\text{C}_6\text{H}_5$ ); the  $^{13}\text{C}\{^1\text{H}\}$  NMR ( $\text{C}_6\text{D}_6^*$ ) has resonances at  $\delta$ : 31.31(d, 13.8Hz,  $\text{CMe}_3$ ), 33.31(d, 24.8Hz,  $\text{CMe}_3$ ), 76.41(d, 2.6Hz,  $\text{C}_5\text{H}_4$ ), 78.58(d, 11.4Hz,  $\text{C}_5\text{H}_4$ ), 85.49(d, 41.6Hz, *ipso*- $\text{C}_5\text{H}_4$ ), 88.38(s,  $\text{C}_5\text{Ph}_5$ ), 126.67(s,  $\text{C}_5\text{Ph}_5$ ), 127.47(s,  $\text{C}_5\text{Ph}_5$ ), 133.23(s,  $\text{C}_5\text{Ph}_5$ ), 136.32(s,  $\text{C}_5\text{Ph}_5$ ); and the  $^{31}\text{P}\{^1\text{H}\}$  NMR ( $\text{C}_6\text{D}_6^*$ ) has a resonance at  $\delta$  25.49(s). Similarly prepared were **P[C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>(C<sub>6</sub>H<sub>4</sub>*p*-CF<sub>3</sub>)<sub>5</sub>](t-butyl)<sub>2</sub>** (37.3% yield, from pentane), **P[C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>(C<sub>6</sub>H<sub>4</sub>*p*-OMe)<sub>5</sub>](t-butyl)<sub>2</sub>** (36% yield, from  $\text{Et}_2\text{O}$  at -30°), **P(C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>HPh<sub>4</sub>)(t-butyl)<sub>2</sub>** (47.4% yield, orange crystals from pentane at -35°), and **P[C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>(3,5-Me<sub>2</sub>C<sub>6</sub>H<sub>3</sub>)<sub>5</sub>](t-butyl)<sub>2</sub>** (30% yield, red solid from pentane/ $\text{Et}_2\text{O}$ ) by using *p*-trifluoromethylchlorobenzene, *p*-methoxychlorobenzene, a limited amount of PhCl and 5-bromo-*m*-xylene respectively. These are air-stable ferrocenyl dialkylphosphines for palladium-catalysed C-C, C-N and C-O bonding-forming cross-coupling reactions. The ligand **P(C<sub>5</sub>H<sub>4</sub>FeC<sub>5</sub>Ph<sub>5</sub>)(t-butyl)<sub>2</sub>** exhibits turnovers of ~1000 for aminations with unactivated aryl bromides or chloride, catalysed the formation of selected aryl ethers under mild conditions, and aryl halides coupled with acyclic or cyclic secondary alkyl- and arylamines, and the latter amines coupled with aryl- and

primary alkylboronic acids. The ligand is air stable in the solid state and in solution, and produces highly active Pd catalysts which are stable in the solid state and react only slowly with oxygen in solution. [Kataoka et al. *J Org Chem* **67** 5553 2002, cf also Shelby et al. *J Am Chem Soc* **122** 10718 2000.]

**4-(3-Phenylpropyl)pyridine 1-oxide (P<sub>3</sub>NO)** [34122-28-6] **M 213.3, m 58-65°**, 60°, **pK<sub>Est</sub>~1.3**. The *N*-oxide is prepared by adding 4-(3-phenylpropyl)pyridine (12.5g, [2057-49-0]) during 5 minutes to a slurry of oxone (37.5g, [70693-62-8]), H<sub>2</sub>O (63.5ml) and MeOH (125ml) (pH ~1.4) under N<sub>2</sub>, and the stirred mixture is kept at pH 5.5 by addition of 5N aqueous NaOH while the temperature is kept at ≤ 35° (the temperature being controlled by external water cooling and by the rate of NaOH addition). The reaction is monitored by HPLC, and at completion, the salts are removed by filtration and the cake is washed with MeOH (50ml). The combined filtrate and washes are treated with 1M aqueous sodium metabisulfite solution (12ml), stirred for 0.5 hours and the pH is adjusted to 10.0 by addition of 5N aqueous NaOH, and set aside for 1 hour. The mixture is concentrated *in vacuo* (at 50°) to a final volume of ~75ml, cooled to 20°, the solid is filtered off, washed with H<sub>2</sub>O (50ml), and dried under N<sub>2</sub> to give the desired *l*-oxide in 93% yield. *Alternatively*, it can be prepared by the oxidation of the *phenylpropylpyridine* with 30% H<sub>2</sub>O<sub>2</sub>/AcOH using the standard method of Ochiai [*J Org Chem* **18** 534 1953, see also Boekelheide & Linn *J Am Chem Soc* **76** 1286 1954]. Its <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has δ at 1.8-1.98 (m, 2H, 2'propyl-CH<sub>2</sub>), 2.3-2.6 (m, 4H, ph-CH<sub>2</sub>, py-CH<sub>2</sub>), 7.05 (m, 2H, 3,5-pyridine), 7.12-7.30 (m, 5H, benzene H), 8.11 (m, 2H, 2,6-pyridine H). [Senanayake et al. *Tetrahedron Lett* **37** 3271 1996.] This reagent stabilises and enhances Jacobsen Mn-salen catalysts and is more readily obtained in quantity than the commonly used 4-phenylpyridine 1-oxide [see Jacobsen et al. *J Am Chem Soc* **113** 7063 1991].

**Pinacolborane (4,4,5,5-tetramethyl-1,3,2-dioxaborolane)** [25015-63-8] **M 128.0, b 42-43°/50mm, d<sub>4</sub><sup>25</sup> 0.882, n<sub>D</sub><sup>20</sup> 1.396**. This borolane is prepared by stirring a cooled solution (at 0°) of pinacol (2.36g, 20mmol [76-09-5]) in CH<sub>2</sub>Cl<sub>2</sub> (2ml) to which is added, dropwise, BH<sub>3</sub>-SMe<sub>2</sub> (20mmol, 10.0M in SMe<sub>2</sub>) whereby effervescence occurs (use efficient fume hood, and gases have a foul odour, are flammable, and should be absorbed through alkaline solution or oxidising solutions containing H<sub>2</sub>O<sub>2</sub> or KMnO<sub>4</sub>). After stirring for ~1 hour at 0°, the mixture is warmed to ~25°, when evolution ceases (*ca* another 1 hour) and is distilled to give the borolane as a clear oil (1.61g, 63% yield). Its IR (CH<sub>2</sub>Cl<sub>2</sub>) has  $\nu_{\max}$  at 3272s, 2980s, 1530m, 1482s, 1333s, 1143s, 983s, 851s cm<sup>-1</sup>; and the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has δ at 1.27 (s, 12H); the <sup>13</sup>C NMR (75.5MHz, CDCl<sub>3</sub>) has δ at 83.1, 24.5; and the <sup>11</sup>B NMR (115.5MHz, CDCl<sub>3</sub>) has δ at 28.07. It is a very efficient hydroboration reagent, requires mild reaction conditions, has higher functional group tolerance, higher regio- and stereoselectivity, with excellent stability of the resulting pinacol boronic esters produced. [Tucker et al. *J Org Chem* **57** 3482 1992.]

**Rhodium(I) chloride 1,5-cyclooctadiene complex dimer {chloro(1,5-cyclooctadiene)rhodium dimer, di-μ,μ'-chlorobis[1,2,5,6-η]-1,5-cyclooctadiene}dirhodium (I), [Rh(COD)Cl]<sub>2</sub>}** [12092-47-6] **M 493.1, m ~243° (dec, see below)**. The complex is prepared by refluxing a solution of RhCl<sub>3</sub>·3H<sub>2</sub>O (1.93g, 7.3mmol), H<sub>2</sub>O (3ml), EtOH (35ml) and COD (6ml, 48mmol) overnight. After cooling, the crude solid is collected, washed with a little H<sub>2</sub>O, dried and recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/hexane or acetic acid, then dried *in vacuo* to give 82% yield (1.5g) of dimer. On heating it darkens from about 220°, melts at ~256° and decomposes with effervescence at 258°. **M** by ebullioscopy in 0.9% CHCl<sub>3</sub> is ~513 confirming the dimeric structure. It is soluble in CH<sub>2</sub>Cl<sub>2</sub> and THF, moderately in CHCl<sub>3</sub>, AcOH and Me<sub>2</sub>CO, slightly soluble in hexane, Et<sub>2</sub>O, MeOH, EtOH, and \*C<sub>6</sub>H<sub>6</sub> but insoluble in H<sub>2</sub>O. Its magnetic susceptibility  $\chi$ , is -0.52 x 10<sup>-6</sup> ±4% per g. [Chatt & Venanzi *J Chem Soc* 4735 1957, Schenck et al. *Inorg Chem* **24** 2334 1985, cf Giordano & Crabtree *Inorg Synth* **19** 218 1979.]

**Rhodium(I) chloride norbornadiene complex dimer {bicyclo[2.2.1]hepta-2,5-diene-rhodium(I) chloride dimer, di-μ,μ'-chlorobis[2,5-norbornadiene]dirhodium(I), [Rh(NBD)Cl]<sub>2</sub>}** [dimer 12257-42-0; polymer 42740-82-9] **M 460.99, m ~240° (dec)**. The yellow dimer is obtained by shaking rhodium "trichloride" (0.7g, RhCl<sub>3</sub>·3H<sub>2</sub>O) and norbornadiene (2ml) in aqueous EtOH (10ml) for 2 days. It is collected and recrystallised from hot CHCl<sub>3</sub>/light petroleum to give analytically pure complex (0.62g). **M** by ebullioscopy in \*C<sub>6</sub>H<sub>6</sub> is 481 consistent with a dimeric structure. The fine yellow crystals are soluble in CHCl<sub>3</sub> and \*C<sub>6</sub>H<sub>6</sub>, but almost insoluble in light petroleum and Et<sub>2</sub>O. Its IR (CS<sub>2</sub>) has  $\nu_{\max}$  at 3060(m), 3000(m), 2960(m), 2920(m), 2855(m), 1307(s), 1171(m), 1157(w), 1068(w), 1029(w), 995(w), 932(m), 882(m), 721(w), 680(s), 630(s) cm<sup>-1</sup>. [Abel,

Bennett & Wilkinson *J Chem Soc* 3178 1959.]

**Ruthenium(II)benzenedichloride dimer {di- $\mu$ -chloro-bis[( $\eta$ -benzene)chlororuthenium(II)],  $\text{Ru}_2(\text{C}_6\text{H}_6)_2\text{Cl}_4$ ,  $[\text{Ru}(\text{C}_6\text{H}_6)\text{Cl}_2]_2$ } [37366-09-9] M 500.2, m 242°.** This versatile reagent is prepared by refluxing  $\text{RuCl}_3 \cdot 3\text{H}_2\text{O}$  (2.0g) in EtOH (100ml) with cyclohexadiene (10ml, of 1,3- or 1-4 prepared by reduction of \*benzene with Na/Liquid  $\text{NH}_3$ , Kaiser *Synthesis* 391 1972, Birch & Subba Rao *Adv Org Chem* 8 1 1972) for 4 hours. The brown precipitate is filtered off, washed with MeOH and dried *in vacuo* to give the complex (1.83g, 95%). The complex is diamagnetic, it is *dimeric* in freezing  $\text{CHBr}_3$ , and is sparingly soluble in organic solvents except  $\text{Me}_2\text{SO}$  with which it forms a red complex with the *monomeric* structure  $\text{RuCl}_2(\text{C}_6\text{H}_6)(\text{dmsO})$ . It dissolves in pyridine, tertiary phosphines, phosphites, tertiary arsines (L) which cleave the chlorine bridges of the dimer to form orange or red air stable complexes  $\text{RuCl}_2(\text{C}_6\text{H}_6)(\text{L})$  which are *monomeric* in  $\text{CHCl}_3$ . Large excess of the ligand L should be avoided in the reaction, otherwise the benzene rings may be displaced. Its IR has Ru—Cl stretching bands with  $\nu_{\text{max}}$  at 295, 259 and 248sh  $\text{cm}^{-1}$ ; and the  $^1\text{H}$  NMR ( $\text{CDCl}_3$ , TMS) has a singlet at  $\tau_{\text{H}}$  4.02 which is *ca* 2.0 ppm upfield of the free benzene H resonance. It is a useful precursor for the preparation of Ru(II) complexes, some of which are catalysts. [Bennett & Smith *J Chem Soc, Dalton Trans* 233 1974, Bennett et al. in *Comprehensive Organometallic Chemistry* Wilkinson ed, Vol 4 pp 748-750 Pergamon Press Oxford 1982.]

The complex  $\text{Ru}_2(\text{C}_6\text{H}_6)_2$  has also been prepared by evaporating a coat of Ru powder (10-15%) in slow curing "Araldite" cement on a 1mm tungsten wire heated to  $\sim 2400^\circ$ . The Ru vapour thus produced is co-condensed with \*benzene to form thermally unstable orange-yellow crystals of  $\text{Ru}_2(\text{C}_6\text{H}_6)_2$  which decompose below  $0^\circ$ . The  $^1\text{H}$  NMR ( $\text{CFCl}_3$ ) of the orange crystals show H resonances at  $\tau$  4.22 (2H), 4.84 (6H), 5.10 (2H) and 7.11(2H) consistent with the 1—4- $\eta$ , 1—6- $\eta$  formulation of the  $\text{Ru}_2(\text{C}_6\text{H}_6)_2$  structure. [Timms & King *J C S, Chem Commun* 898 1978.]

**Ruthenium dichloride 1,5-cyclooctadiene complex [dichloro- $\mu$ -( $\eta^4$ -cycloocta-1,5-diene)ruthenium(II),  $\text{RuCl}_2(\text{COD})_n$ ] [50982-12-2] M 280.2 (monomer).** Analytically pure complex is obtained by heating a mixture of pure  $\text{RuCl}_3 \cdot 3\text{H}_2\text{O}$  (50.0g, 0.2mol) and cycloocta-1,5-diene (50ml, 0.4mol, use a fume hood) in absolute EtOH (400ml) under reflux for 24 hours whereby the brown complex precipitates. After cooling, the solid is filtered off, washed with  $\text{Et}_2\text{O}$  ((50ml) and dried *in vacuo* to give pure  $[\text{RuCl}_2(\text{COD})]$  ( $\sim 52\text{g}$ ,  $\sim 90\%$ ). It is stable in air, only slightly soluble in organic solvents, reacts by halogen-bridge cleavage and is a useful synthetic precursor of ruthenium(II) complexes. [Ohta, Takaya and Noyori *Inorg Chem* 27 566 1988, Alders, Singleton and Yates *Inorg Synth* 26 253 1989, Bennett et al. in *Comprehensive Organometallic Chemistry* Wilkinson ed., Vol 4 pp 748-750 Pergamon Press Oxford 1982.]

**Tetrakis(acetonitrile)copper(I) hexafluorophosphate  $[\text{Cu}(\text{MeCN})_4^+ \text{PF}_6^-]$  [64443-05-6] M 372.7, m 160°(dec).** In an efficient fume cupboard, 60-65% aqueous hexafluorophosphoric acid (10ml,  $\sim 113\text{mmol}$ , [16940-81-1], M 146.0) is added in 2ml portions to a stirred suspension of Cu(I) oxide (4.0g, 28mmol) in MeCN (80ml) when a very exothermic reaction occurs (use a reflux condenser), causing the solution to boil. After stirring for a further 3 minutes the solution is filtered (use a medium-porosity frit) from some black solid, and  $\text{Cu}(\text{MeCN})_4^+ \text{PF}_6^-$  may begin to crystallise out, and is complete on standing at  $-20^\circ$ , or by adding an equal volume of dry  $\text{Et}_2\text{O}$  and cooling to  $0^\circ$ . (If some white solid separates with the black solid it should be washed off with MeCN and the MeCN solutions combined and cooled at  $-20^\circ$ , or diluted with  $\text{Et}_2\text{O}$ .) The salt is collected by filtration, washed with  $\text{Et}_2\text{O}$  and immediately dissolved in MeCN (100ml), filter off any undissolved blue  $\text{Cu}^{2+}$  species, add  $\text{Et}_2\text{O}$  (100ml) and set aside at  $-20^\circ$  for several hours to crystallise out. The crystalline complex would need a second recrystallisation if it has a blue tinge. It may be recrystallised from MeCN (80ml) and  $\text{Et}_2\text{O}$  (80ml). The white complex salt is dried *in vacuo* for  $\sim 30$  minutes immediately after washing with  $\text{Et}_2\text{O}$  to give the analytically pure salt ( $\sim 12.5\text{g}$ ,  $\sim 60\%$  depending on crystallisation losses). It is a free-flowing white, microcrystalline powder which does not darken on long storage under  $\text{N}_2$  or argon. On exposure to air for  $>1$  hour minor surface oxidation occurs as it is slightly hygroscopic. The MeCN ligands are not removed easily *in vacuo* at  $\sim 25^\circ$  (but dissociation occurs at  $\sim 5$  torr/ $80^\circ$  and  $\sim 25$  torr/ $110^\circ$ ). The IR (Nujol) has  $\nu_{\text{max}}$  at 2277m and 2305m  $\text{cm}^{-1}$  (for MeCN), and 850vs and 557s  $\text{cm}^{-1}$  for  $\text{PF}_6^-$ . The related salt,

$\text{Cu}(\text{MeCN})_4^+ \text{BF}_4^-$ , can be prepared similarly by using an equivalent amount of  $\text{HBF}_4$  (as 48% aqueous acid [16872-11-0], M 87.9) instead of  $\text{HPF}_6$ . These complexes are suited for the non-aqueous-media synthesis of Cu(I) complexes [see  $(\text{IPr})_2\text{Cu}^+ \text{PF}_6^-$  and  $(\text{IPr})_2\text{Cu}^+ \text{BF}_4^-$  above]; as well as for preparing other complexes such as  $[\text{Rh}(\text{C}_2\text{H}_4)_3(\text{MeCN})_2]^+ + \text{CuCl}$  by reaction of  $\text{Cu}(\text{MeCN})_4^+$  with  $[\text{Rh}(\text{C}_2\text{H}_4)_2\text{Cl}]_2$  and  $\text{C}_2\text{H}_4$  in  $\text{CH}_2\text{Cl}_2$  [Maspero et al. *J Organomet Chem* **38** C43 1972]. The X-ray structure of the  $\text{ClO}_4^-$  salt showed that the copper is almost ideally tetrahedral with nearly linear MeCN molecules [Csoregh et al. *Acta Cryst* **B31** 314 1975]. [Kubas *Inorg Synth* **19** 90 1979, Kubas *Inorg Synth* **28** 68 1990.]

**cis-cis-cis-1,2,3,4-Tetrakis(diphenylphosphinomethyl)cyclopentane** [(1*R*\*,2*R*\*,3*S*\*,4*S*\*)-tetrakis(diphenylphosphinomethyl)cyclopentane, **Tedicyp**] [333380-86-2] **M 862.9, m 790**. The intermediate cis-cis-cis-1,2,3,4-tetrakis(boranatodiphenylphosphanyl)methyl)cyclopropane (white crystals **m 1490**, after chromatography on silica gel eluting with AcOEt/petroleum ether, 1:4) is obtained from  $\text{Ph}_2\text{PLi}$  (from  $\text{Ph}_2\text{PCl}$  see [1079-66-9] and Li) and cis-cis-cis-1,2,3,4-tetrakis(tosylloxymethyl)cyclopentane in THF under argon (~20°/5 hours), then borane in THF is added. This boranato intermediate formed is added to anhydrous  $\text{Et}_2\text{NH}$ , heated at 55-60° for 10 hours, evaporated to dryness, this process is repeated twice, and the residue is chromatographed on silica gel (eluted with  $\text{Et}_2\text{O}$ /petroleum ether, 1:4) to give white crystals of **Tedicyp** (98%) with **m 790**. It is not very stable in air and must be stored under argon. It has  $^1\text{H}$  NMR (400MHz, THF- $d_8$ ) with  $\delta$  at 7.71-7.23 (40H, m), 2.91-2.52 (2H, m), 2.38-1.72 (10H, m), 1.25 (2H, m); the  $^{13}\text{C}$  NMR (100MHz, THF- $d_8$ , TMS) has  $\delta$  at 141.6 (s,  $^1J_{\text{P,C}} = 15.1\text{Hz}$ ), 141.1 (s,  $^1J_{\text{P,C}} = 14.1\text{Hz}$ ), 140.0 (s,  $^1J_{\text{P,C}} = 15.1\text{Hz}$ ), 139.4 (s,  $^1J_{\text{P,C}} = 16.0\text{Hz}$ ), 134.6 (d,  $J_{\text{P,C}} = 20.1\text{Hz}$ ), 134.5 (d,  $J_{\text{P,C}} = 19.1\text{Hz}$ ), 133.5 (d,  $J_{\text{P,C}} = 17.0\text{Hz}$ ), 133.3 (d,  $J_{\text{P,C}} = 18.0\text{Hz}$ ), 129.6 (d,  $J_{\text{P,C}} = 14.0\text{Hz}$ ), 129.3-129.0 (d,  $J_{\text{P,C}} = 15.0\text{Hz}$ ), 45.4 (t,  $^2J_{\text{P,C}} = 8.6\text{Hz}$ ), 45.2 (d,  $^2J_{\text{P,C}} = 8.7\text{Hz}$ ), 40.4 (t,  $^2J_{\text{P,C}} = 8.75\text{Hz}$ ), 40.3 (t,  $^2J_{\text{P,C}} = 9.2\text{Hz}$ ), 3.91 (d,  $^2J_{\text{P,C}} = 13.3\text{Hz}$ ), 33.1 (d,  $^1J_{\text{P,C}} = 10\text{Hz}$ ), 26.5 (t,  $^1J_{\text{P,C}} = 12.9\text{Hz}$ ); the  $^{31}\text{P}$  NMR (162MHz, THF- $d_8$ ) has  $\delta$  at -16.3 and -17.7. It is used for preparing the Pd-Tedicyp catalyst with  $[\text{Pd}(\eta^3\text{-C}_3\text{H}_5)\text{Cl}]_2$  (see above) for promoting allylic alkylations [Laurenti et al. *J Org Chem* **66** 1633 2001], cross coupling reactions [Wallow & Novak *J Org Chem* **59** 5034 1994], and catalysing Heck reactions of  $\alpha$ - or  $\beta$ -substituted enol ethers with aryl bromides [Battace et al. *Eur J Org Chem* 3122 2007] with high efficiency.

The tetra-phosphine oxide of cis-cis-cis-1,2,3,4-tetrakis(diphenylphosphino)methyl)cyclopropane is obtained by oxidising **Tedicyp** in THF with excess  $\text{H}_2\text{O}_2$  (~20°/4 days) then purified by chromatography on silica gel (eluting with MeOH/ $\text{CH}_2\text{Cl}_2$ , 5:95) to yield (66%) white crystals of the *tetroxide* with **m 1650**. It has  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) with  $\delta$  at 7.79-7.62 (16H, m), 7.32-7.28 (8H, m), 2.85 (2H, br t,  $J = 12.9\text{Hz}$ ), 2.28 (4H, d,  $J = 11.9\text{Hz}$ ), 2.26-2.15 (2H, m), 2.10-2.05 (2H, m), 1.88 (2H, br q,  $J = 12.9\text{Hz}$ ), 1.60 (1H, dt,  $J = 13.8, 7.5\text{Hz}$ ), 1.36 (1H, dt,  $J = 13.8, 10.2\text{Hz}$ ); the  $^{13}\text{C}$  NMR (100MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 134.3 ( $^1J_{\text{P,C}} = 13.0\text{Hz}$ ), 133.3 (m), 131.94 ( $^1J_{\text{P,C}} = 5\text{Hz}$ ), 131.62 ( $^1J_{\text{P,C}} = 5\text{Hz}$ ), 131.27-131.0 (m), 129.0-128.6 (m), 41.9 (d,  $^2J_{\text{P,C}} = 10.0\text{Hz}$ ), 38.0 (t), 35.2 (d), 33.0 (t,  $^1J_{\text{P,C}} = 69.9\text{Hz}$ ), 27.1 (t,  $^1J_{\text{P,C}} = 70.0\text{Hz}$ ) relative to TMS; and the  $^{31}\text{P}$  NMR (162MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 32.2. [Laurenti et al. *J Org Chem* **66** 1633 2001.]

**R-(+)- and S-(-)- 2,2',6,6'-Tetramethoxy-4,4'-bis(diphenylphosphino)-3,3'-bipyridine** (**R-(+)- P-Phos** and **S-(-)- P-Phos**) [*R*- 221012-82-4 and *S*- 362524-23-0] **M 644.6, m 261-2650**,  $[\alpha]_{\text{D}}^{20} +980$  and  $-980$  (c 1,  $\text{CHCl}_3$ ). This group of ligands have been prepared from (3-bromo-2,6-dimethoxy-pyrid-4-yl)(diphenyl)-phosphine oxide by an Ulmann reaction ( $\text{Cu}/\text{DMF}/140^\circ$ ) to give *RS*- 2,2',6,6'-tetramethoxy-3,3'-bipyridine-4,4'-diphenylphosphine dioxide (85%) which is resolved into the (+)- and (-)- enantiomers by HPLC on a preparative Diacel AD column (25m x 1.2m), eluting with 2-propanol/hexane (20:80) at a flow rate of 3.0ml/minute:  $t_{\text{R}} = 12.24$  minutes,  $t_{\text{S}} = 25.06$  minutes. The phosphine oxide groups are then each reduced by  $\text{Cl}_3\text{SiH}/\text{Et}_3\text{N}/\text{toluene}$  at  $140^\circ$  to give *R-(+)-* and *S-(-)-* **R-Phos** in 99% yields. Alternatively, the above *RS*- dioxide, which cannot be resolved easily, is converted into the 5,5'-dibromo derivative ( $\text{Br}_2/\text{NaOAc}/\text{AcOH}$ ;  $0^\circ$ - $60^\circ$ , 97%) which can now be resolved via the (+) and (-)-di-*O*-benzoyl tartrate salts (in  $\text{EtOAc}/\text{CHCl}_3$ , 1/1) to provide the enantiomeric 5,5'-dibromo-dioxides that are de-brominated, and reduced ( $\text{Cl}_3\text{SiH}/\text{Et}_3\text{N}/\text{toluene}$  during an extended period from  $25^\circ$  to  $140^\circ$ ) to give **R-(+)- and S-(-)- P-Phos** in 91% yields. [Pai et al. *J Am Chem Soc* **112** 11513 2000.]

**R-(+)- and S-(-)- 2,2',6,6'-Tetramethoxy-4,4'-bis(di(3,5-xylyl)phosphino)-3,3'-bipyridine** [**R-(+)- Xylyl-P-Phos** and **S-(-)- Xylyl-P-Phos**] [*R*- 442905-33-1 and *S*- 443347-10-2] **M 756.9, m (158-1620), 190-1940**,  $[\alpha]_{\text{D}}^{20} +1250$  and  $-1250$  (c 1,  $\text{CHCl}_3$ ). The Xylyl-P-Phos ligands are prepared in a slightly different way from P-Phos ligands above. Thus 3-bromo-2,6-dimethoxy-pyridine is converted into (3-bromo-2,6-dimethoxy-pyrid-4-yl)di(3,5-dimethylphenyl)phosphine [*a*:  $\text{LDA}/\text{THF}/-78^\circ$ ; *b*: (3,5-dimethylphenyl) $_2\text{PCl}/-78^\circ$ ] in 56% yield, then

oxidised ( $\text{H}_2\text{O}_2/\text{Me}_2\text{CO}/0^\circ$ ) to the *phosphine oxide* (99%), followed by an Ulmann reaction as above ( $\text{Cu}/\text{DMF}/140^\circ$ ) to give *RS-2,2',6,6'-tetramethoxy-3,3'-bipyridine-4,4'-di(3,5-dimethylphenyl)phosphine dioxide* (85% yield) that is resolved *via* (-)- and (+)- di-*O*-benzoyl tartrate salts (*a*: see above fractional crystallisation; *b*: 10% aqueous NaOH) to give the *R*- (75%) and *S*- (88%) *Xylyl-P-Phos dioxides* which are reduced ( $\text{Cl}_3\text{SiH}/\text{Et}_3\text{N}/\text{toluene}$ ) to *R*(+)- and *S*- *Xylyl-P-Phos* in >90% yields. The absolute configuration *S*- was determined from the X-ray crystal structure of the diastereoisometrically pure 1:1 salt of (-)-*Xylyl-P-Phos* and (+)-(*2S,3S*)-dibenzoyl-*O*-tartrate. The latter have  $^1\text{H}$  NMR (500MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{H}}$  at 2.20 (s, 12H,  $\text{PhCH}_3$ ), 2.25 (s, 12H,  $\text{PhCH}_3$ ), 3.37 (s, 6H, O  $\text{CH}_3$ ), 3.83 (s, 6H, O  $\text{CH}_3$ ), 6.06 (d,  $J = 1.5\text{Hz}$ , 2H,  $\text{PyH}$ ), 6.79-6.92 (m, 12H,  $\text{PhH}$ ); the  $^{13}\text{C}$  NMR (125MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{C}}$  at 21.54 (d,  $J = 4.78\text{Hz}$ ), 53.06, 53.43, 105.47, 115.39 (t,  $J = 37.07\text{Hz}$ ), 130.47 (d,  $J = 33.81\text{Hz}$ ), 131.46 (t,  $J = 20.99\text{Hz}$ ), 132.42 (t,  $J = 21.74\text{Hz}$ ), 135.69 (t,  $J = 10.05\text{Hz}$ ), 136.88 (t,  $J = 12.82\text{Hz}$ ), 137.39 (t,  $J = 8.04\text{Hz}$ ), 137.71 (t,  $J = 7.16\text{Hz}$ ), 154.64, 154.70, 154.75, 160.76 (t,  $J = 11.18\text{Hz}$ ), 162.30; the  $^{31}\text{P}$  NMR (200MHz,  $\text{CDCl}_3$ ) with  $\delta_{\text{P}}$  at -11.99; the LSMS [ $\text{M}^+$ ] found 757, calculated for  $\text{C}_{46}\text{H}_{50}\text{N}_2\text{O}_4\text{P}_2$  is 756.85; and the C, H and N elemental analyses fit for the latter empirical formula. [Wu et al. *Tetrahedron Lett* **43** 1539 2002.]

**6RS(±)-4,4,6-Trimethyl-2-vinyl-1,3,2-dioxaborinane (vinylboronic acid 2-methyl-2,4-pentanediol ester)** [4627-10-5] **M 154.0, b 50-55°/0.46mm,  $d_4^{25}$  0.893,  $n_D^{20}$  1.429**. It is prepared by a modification of a published method [Hoffmann & Landmann *Chem Ber* **119** 2013 1986] where *n*-octanol is replaced by 2-methylpentane-2,4-diol (MPD). To a solution of  $(\text{MeO})_3\text{B}$  (30ml, 270mmol) in dry  $\text{Et}_2\text{O}$  (125ml) at  $-78^\circ$  is added over 45 minutes a 2.1M solution of vinylmagnesium chloride (118ml, 248mmol) in THF with stirring, and allowing the temperature to rise to  $\sim 25^\circ$ . The mixture is then acidified, under cooling, with aqueous HCl (concentrated, 21ml, 0.25mol in 62.5ml of  $\text{H}_2\text{O}$ ), phenothiazine (0.1g, as stabiliser) is added, the phases are separated, the aqueous phase is extracted with MPD (3 x 50ml). The combined organic phases are evaporated *in vacuo* from a bath at  $80^\circ$  and the residual oil is distilled at high vacuum to give pure *borinane* in  $\sim 60\%$  yield. [Note that all the liquids in the reaction mixture distil at lower temperatures than the desired racemic vinyl-dioxaborinane.] It is superior to vinylboronic pinacol ester in preparation, storage, stability, and reactivity, providing improved selectivity for Heck *versus* Suzuki coupling with aryl and heteroaryl bromides and iodides [Lightfoot et al. *Tetrahedron Lett* **44** 7645 2003]. By adopting different reaction conditions with  $\text{Pd}(\text{PPh}_3)_4/t\text{-BuOK}$ , selective Suzuki-Miyaura coupling with a range of aryl and heteroaryl halides (Cl, Br and I) was achieved [Lightfoot et al. *Synlett* 259 2005, cf Lightfoot et al. *Org Biomol Chem* **3** 3167 2005].

**Tri-*tert*-butylphosphonium tetrafluoroborate [(*tert*-Bu) $_3\text{PH}^+\text{BF}_4^-$ ] [131274-22-1] **M 290.1, m 261°(dec),  $\text{pK}^{25}$  11.4 (phosphine)**. The salt is prepared by adding  $\text{HBF}_4$  (48% aqueous solution, 1.0ml, 7.6mmol) to a solution of pure (*tert*-Bu) $_3\text{P}$  (225mg, 1.1mmol, see above) in  $\text{CH}_2\text{Cl}_2$  (15ml), stirring vigorously for 5 minutes, the organic layer is separated, dried ( $\text{MgSO}_4$ ), filtered and evaporated to dryness to give analytically pure *salt* (302mg, 94%) as a white powder. It has  $\nu_{\text{max}}$  (KBr) at 3002, 2217, 1478, 1405, 1381, 1181, 1059, 886 (br)  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (400MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{H}}$  at 6.07 (d,  $^1J_{\text{PH}} = 465\text{Hz}$ , 1H), 1.65 (d,  $^3J_{\text{PH}} = 15.3\text{Hz}$ , 27H); the  $^{13}\text{C}$  NMR (75MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{C}}$  at 37.0 (d,  $^1J_{\text{PC}} = 29\text{Hz}$ , 1H), 30.1; and the  $^{31}\text{P}\{^1\text{H}\}$  NMR (121MHz,  $\text{CDCl}_3$ ) has  $\delta_{\text{P}}$  at 51.7. The salt is more air-stable than the free phosphine, and can be heated at  $120^\circ$  for 24 hours without decomposition (no change in NMR spectra), or loss of catalytic activity with transition metal complexes. It is *not hygroscopic*. In the presence of  $\text{Pd}_2(\text{dba})_3$  in THF it catalyses Suzuki cross-coupling reactions between aryl halides and arylboronic acids, Stille cross-coupling reactions between aryl halides and tributylSn compounds, and Heck reactions between aryl halides and olefines [Netherton & Fu *Org Lett* **3** 4295 2001]. The salt has also been used successfully in the Heck coupling [with  $\text{PdCl}_2(\text{COD})$ , LiCl,  $\text{C}_2\text{N}_2\text{Me}$ ] of non-activated alkenyl tosylates and phosphates as substrates with electron-poor alkenes and styrene derivatives [Hansen et al. *Angew Chem Int Ed* **45** 3349 2006].**

**Vinyl MIDA boronate {vinylboronic acid MIDA ester, 6-methyl-2-vinyl-1,3,6,2-dioxazaborocane-4,8-dione, [*N*-[(*carboxy-κO*)methyl]-*N*-methylglycinato(2-)-κN,κO]ethenyl boron}** [1104636-73-8] **M 183.0, m 152-156°**. The boronate is prepared in Schlenk equipment by adding dropwise vinyltrimethylsilane (4.49ml, 31.5mmol, freshly distilled see [754-05-2]) to a stirred solution of  $\text{BBr}_3$  (1.0M in  $\text{CH}_2\text{Cl}_2$ , 30mmol) at  $0^\circ$ , and maintained at this temperature for 20 minutes, then allowed to warm to  $\sim 25^\circ$  with stirring for a further 2 hours. This mixture is added *via* a cannula to a stirred suspension of MIDA sodium salt (5.73g, 30.0mmol, see above)

in MeCN (50ml) at 0°, at such a rate as to keep the temperature below 5°; then the temperature is allowed to warm to ~25° while stirring for 1 hour. The resulting white suspension is filtered through a pad of Celite and the filtrate cake is extracted 3 times with Me<sub>2</sub>CO. Et<sub>2</sub>O is added to the combined orange filtrates which allowed the colourless free flowing vinyl MIDA boronate (4.74g, 86%) to crystallise out. On TLC it has R<sub>F</sub> 0.26 (Merck silica gel plate grade 9385, 60Å, 230-400 mesh, with EtOAc). Its IR (thin film) has  $\nu_{\max}$  at 3063, 2997, 2960, 1755, 1617, 1455, 1420, 1345, 1312, 1251, 1175, 1155, 1134, 1117, 1090, 1033, 987, 964, 951, 865 cm<sup>-1</sup>; the <sup>1</sup>H NMR (500MHz, Me<sub>2</sub>CO-*d*<sub>6</sub>,  $\delta$  = 2.04 centre line) has  $\delta_{\text{H}}$  at 5.96 (dd,  $J$  = 19.0, 13.5Hz, 1H), 5.72-5.63 (m, 2H), 4.21 (d,  $J$  = 17.0Hz, 2H), 4.01 (d,  $J$  = 17.0Hz, 2H), 3.0 (s, 3H); the <sup>13</sup>C NMR (125MHz, Me<sub>2</sub>CO-*d*<sub>6</sub>,  $\delta$  = 29.80 centre line) has  $\delta_{\text{C}}$  at 168.3, 128.7, 61.6, 46.7; and the HMRS (EI+) has found  $m/z$  183.0700, calculated for C<sub>7</sub>H<sub>10</sub>BBrNO<sub>4</sub> [M]<sup>+</sup> is 183.0703.

Single crystal X-ray analysis confirmed the predicted structure as having a pyramidalised boron centre. The vinylboronate shows no signs of deterioration when kept on the benchtop in air for more than 3 months. It is a versatile reagent that can be prepared on a multigram scale. It readily reacts with CH<sub>2</sub>N<sub>2</sub>/Pd(OAc)<sub>2</sub> to provide cyclopropyl MIDA boronate [Et<sub>2</sub>O, 0-23°, 1 hour, 93%], and with *m*-CPBA it yields oxiranyl MIDA boronate [CH<sub>2</sub>Cl<sub>2</sub>, 0-23°, 18 hours, 74%] [Uno, Gillis and Burke *Tetrahedron* **65** 3230 2009]; and can be used in Suzuki, Heck, Stille, Negishi and Sonogashira couplings and Miyaura borylations [Gillis & Burke *Aldrichimica Acta* **42** 17 2009].

**Xantphos [4,5-Bis(diphenylphosphino)-9,9-dimethylxanthene]** [161265-03-8] **M 578.2, m 221-222°, 224-228°**. Xantphos is prepared, using Schlenck techniques under N<sub>2</sub>, by adding *sec*-BuLi (13.6ml of 1.3M in 98/2 cyclohexane/hexane, 14.3mmol) dropwise to a stirred solution of 9,9-dimethylxanthene (1.0g, 4.76mmol, see [19814-75-6]) and TMEDA (2.3ml, 13.3mmol) in dry Et<sub>2</sub>O (54ml), and stirring for 16 hours at ~25° to form the 4,5-dilithium derivative. To this mixture is added a solution of chlorodiphenylphosphine (2.8ml, 14.3mmol, see [1079-66-9]) in hexane (16ml) dropwise, and stirring is continued for 16 hours, then the solvent is removed *in vacuo*. The residual thick oil is dissolved in CH<sub>2</sub>Cl<sub>2</sub>, washed with H<sub>2</sub>O, dried (MgSO<sub>4</sub>), filtered, the solvent is removed *in vacuo*, the residue is washed with hexanes, and recrystallised from propan-1-ol to give *xantphos* (2.05g, 74.6%) as an air stable yellow-white powder. Its IR (CHCl<sub>3</sub>) has  $\nu_{\max}$  at 3.073w, 2974w, 1435s, 1405vs, 1243m, 695m cm<sup>-1</sup>; the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 7.4 (dd, 2H,  $J$  = 7.8, 1.0Hz, CPCHCH), 7.15-7.26 (aryl, 20H, [P(C<sub>6</sub>H<sub>5</sub>)<sub>2</sub>]<sub>2</sub>), 6.96 (t, 2H,  $J$  = 7.7Hz, CHCHCH), 6.54 (dd, 2H,  $J$  = 7.4, 1.4Hz, CHCHCC), 1.65 (s, 6H, CH<sub>3</sub>); the <sup>13</sup>C{<sup>1</sup>H}NMR (100MHz, CDCl<sub>3</sub>) has  $\delta$  at 137.2 (t,  $J$  = 5.3Hz) through space P-P coupling ~60Hz, 133.7 (t,  $J$  = 10.1Hz), 131.9, 129.7, 128.0 (ar), 126.1, 125.7 (t,  $J$  = 9.8Hz), 123.1, 67.8 (CMe<sub>2</sub>), 31.6 (CH<sub>3</sub>); the <sup>31</sup>P{<sup>1</sup>H}NMR (121.5MHz, CDCl<sub>3</sub>, referenced to external 85% H<sub>3</sub>PO<sub>4</sub>) has  $\delta$  at -17.5; and exact mass (MS) is found to be 578.1916 (calculated for C<sub>39</sub>H<sub>32</sub>OP<sub>2</sub> is 578.1928). Out of several diphenylphosphino ligands studied, *xantphos* induced the highest selectivity in rhodium-catalysed hydroformylation of 1-alkenes to form linear aldehydes. The hydroformylation is carried out in a stainless steel bomb under pressure of ~10 bars of CO/H<sub>2</sub> (1:1) in toluene with phosphorus ligand and Rh(acac)(CO)<sub>2</sub> as metal source. [Kranenburg, van der Burgt, Kramer and van Leewen *Organometallics* **14** 3081 1995.]

## PHASE TRANSFER CATALYSTS

Phase transfer catalysts (PTCs) are agents that are soluble in both lipophilic and hydrophilic solutions, i.e. soluble in aqueous solutions and in non-water soluble organic solvents, in varying degrees. They are mostly cationic species that assist the transfer of anions. e.g. CN<sup>-</sup>, from the aqueous phase into the non-aqueous phase, e.g. hexane, CHCl<sub>3</sub>, which would contain the reactive compound, e.g. 1-octyl chloride, to form the product, i.e. 1-octyl nitrile, in the non-aqueous phase in high yields. However, the catalysts can also be essentially neutral as with *crown ethers*, *cryptands* (tricyclic ethers containing nitrogen atoms at the bridgehead allowing formation of three rings) and *ionophores* which are useful for transferring metal ions through the bilayer interface. Some PTCs are polymer-bound making it easy for their removal and recycling. Other reactions that are phase transfer catalysed include:  $\beta$ -eliminations, *N*-alkylations, nucleophilic aromatic substitutions, metal-organic reactions, oxidation and reduction reactions. The reactions may proceed in either phase and/or at the interphase of the phases. They invariably occur in the organic phase, but in *Reverse Phase Transfer* reactions they occur in the aqueous phase. This section provides a selection of such agents, some of which are chiral and are capable of

promoting stereoselective, or stereospecific reactions. [Starks *J Am Chem Soc* **93** 195 1971, Brändström *Adv Phys Org Chem* **15** 267 1977, G.W. Gokel *Crown Ethers and Cryptands*, Royal Society of Chemistry, Cambridge, 1991, ISBN 0-85186-996-3; S. R. Cooper (Ed.), *Crown Compound: Towards Future Applications*, VCH, Weinheim, 1992, ISBN 1-56081-024-6, 3-527-28073-1; B. Deitrich, P. Viout and J-M, Lehn, *Macrocyclic Chemistry: Aspects of Organic and Inorganic Supramolecular Chemistry*, VCH, Weinheim, 1993, ISBN 3-527-28330-7; M. Hiraoka (Ed.), *Crown Ethers and Analogous Compounds*, Elsevier, Amsterdam, 1992, ISBN 0-444-88191-3; Ooi & Maruoka *Aldrichimica Acta* **40** 77 2007; K. Maruoka (Ed.), *Asymmetric Phase Transfer Catalysis*, Wiley-VCH, Weinheim, 2008, ISBN 978-3-527-31842-1.]

**Adogen 464 [methyl trialkyl(C8—C10)ammonium chloride]** [72749-59-8] **M 416.2, d<sup>25</sup> 0.898, n<sub>D</sub><sup>20</sup> 1.4665**. This is purified in the same way as Aliquat 336 in the following entry. [US Patent 3 992 432.]

**Aliquat 336 (methyltricaprylammonium chloride, tri-*n*-octylmethylammonium chloride)** [5137-55-3, a replacement product, Aliquat 128, has 63393-96-4] **M 404.2, d<sup>25</sup> 0.884, n<sub>D</sub><sup>20</sup> 1.4665**. A 30% (v/v) of Aliquat 336 solution in \*benzene is washed twice with an equal volume of 1.5M HBr. [Petrow & Allen, *Anal Chem* **33** 1303 1961.] It is purified by dissolving 50g in CHCl<sub>3</sub> (100ml) and shaking with 20% NaOH solution (200ml) for 10 minutes, and followed by 20% NaCl (200ml) for 10 minutes. It is then washed with a small volume of H<sub>2</sub>O, and filtered through a dry filter paper. [Starks *J Am Chem Soc* **93** 195 1971, Adam & Pribil *Talanta* **18** 733 1971.]

***N*-Anthracenylmethyl)cinchonidinium chloride** [199588-80-2] **M 521.1, m ~158-161°**. The salt can be purified by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O and dried *in vacuo*. This chiral phase-transfer-catalyst (PTC) assists in the allyl alkylation of the α-carbon of glycine imino ester to yield finally α-allyl-substituted α-amino acids. Thus with a ratio of imino ester to allyl compound: achiral phosphines (e.g. 1,2-bis[diphenylphosphino]ethane DIPHOS, DPPE [1663-45-2], its oxide, triphenoxyphosphine or its oxide) : base (KOH) : [Pd(*p*-allyl)Cl]<sub>2</sub>: PTC of 100:200:150:3.5:10 produced the α-allyl imino ester in varying yields and stereoselectivity. However, the derivative ***N*-(9-anthracenylmethyl)-*O*(9)-methyl-cinchonidinium iodide** (prepared from the chloride with MeI/50% aqueous NaOH/20°/4 hours) was a more effective PTC, providing better yields and with up to ~94% enantiomeric excess. The orange *iodide*, which is soluble in CH<sub>2</sub>Cl<sub>2</sub>, is crystallised from CH<sub>2</sub>Cl<sub>2</sub>/Et<sub>2</sub>O and has [α]<sub>D</sub><sup>19</sup> -288° (c 1, CH<sub>2</sub>Cl<sub>2</sub>), IR with ν<sub>max</sub> (CDCl<sub>3</sub>) at 3686, 3014, 2952, 1598, 1515, 1456, 1075, 986 cm<sup>-1</sup>; the <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) has δ<sub>H</sub> at 1.38-1.55 (m, 2H), 1.95-2.03 (m, 1H), 2.12-2.21 (m, 1H), 2.35-2.43 (m, 1H), 2.64-2.72 (m, 1H), 3.09 (dd, 1H, *J* = 5.6, 5.6Hz), 3.83 (s, 3H), 4.42-4.51 (m, 1H), 4.84 (d, *J* = 4.8Hz, 1H), 5.07 (d, *J* = 5.2Hz, 1H), 5.19 (d, *J* = 8.5Hz, 1H), 5.25 (bs, 1H), 5.87-6.02 (m, 2H), 6.52 (bs, 1H), 6.97 (d, *J* = 6.9Hz, 1H), 7.54-7.60 (m, 3H), 7.69-7.73 (m, 2H), 7.82-7.89 (m, 2H), 7.97 (brs, 1H), 8.05 (d, *J* = 4.3Hz, 1H), 8.14 (d, *J* = 4.1Hz, 1H), 8.20 (d, *J* = 4.3Hz, 1H), 8.32 ((d, *J* = 4.6Hz, 1H), 8.65 (s, 1H), 9.06 (d, *J* = 2.3Hz, 2H), 9.65 (d, *J* = 4.4Hz, 1H); and the <sup>13</sup>C NMR (125MHz, CD<sub>3</sub>OD) has δ<sub>C</sub> at 150.9, 149.1, 143.0, 138.5, 134.7, 134.6, 133.8, 133.0, 132.9, 131.5, 131.2, 131.1, 130.4, 129.6, 129.4, 129.4, 127.2, 126.1, 126.57, 125.7, 124.8, 124.7, 121.8, 118.9, 117.8, 79.5, 70.0, 63.3, 57.7, 57.5, 53.7, 39.4, 27.2, 26.2, 23.2. This *iodide* in CH<sub>2</sub>Cl<sub>2</sub> was converted to another useful PTC ***N*-(9-anthracenylmethyl)-*O*(9)-methyl-cinchonidinium hexafluorophosphate** [by addition of NH<sub>4</sub>PF<sub>6</sub> in H<sub>2</sub>O at 25°/1 hour, dilution with CH<sub>2</sub>Cl<sub>2</sub> and H<sub>2</sub>O, drying the organic layer (MgSO<sub>4</sub>) and evaporating] which gave yellow crystals from CH<sub>2</sub>Cl<sub>2</sub>/CHCl<sub>3</sub> with **m 163-164°**, [α]<sub>D</sub><sup>23</sup> -314° (c 1.1, CH<sub>2</sub>Cl<sub>2</sub>), and IR and NMR quite similar to those of the *iodide*. [Nakoji et al *J Org Chem* **67** 7418 2002.]

**8*S*,9*R*-(-)-*N*-Benzylcinchonidinium chloride (BCDC)** [69257-04-1] **M 421.0, m 212-213° (dec), [α]<sub>D</sub><sup>20</sup> -180°, -183° (c 5, H<sub>2</sub>O), p*K*<sub>Est</sub> ~5**. Dissolve the chloride in the minimum volume of H<sub>2</sub>O and add absolute Me<sub>2</sub>CO. Filter it off and dry it in a vacuum. It can also be recrystallised from hot EtOH or EtOH/Et<sub>2</sub>O. It is a good chiral phase transfer catalyst producing the opposite enantiomer to the one using *N*-benzylcinchoninium chloride as catalyst — see below. [Colonna et al. *J Chem Soc, Perkin Trans 1* 547 1981, Imperali & Fisher *J Org Chem* **57** 757 1992.] [Beilstein **23** H 446.] See cinchonidine [485-71-2].

**8*R*,9*S*-(+)-*N*-Benzylcinchoninium chloride (BCNC, *N*-benzyl-9*S*-hydroxycinchoninium chloride)** [69221-14-3] **M 421.0, m 265° (dec), [α]<sub>D</sub><sup>20</sup> +169° (c 0.4, H<sub>2</sub>O), p*K*<sub>Est</sub> ~ 5**. Recrystallise the chloride from isoPrOH, toluene or small volumes of H<sub>2</sub>O. It is a good chiral phase transfer catalyst producing the opposite enantiomer to

the one using the above *N*-benzylcinchonidinium chloride as catalyst — see above. [Julia et al. *J Chem Soc, Perkin Trans 1* 574 1981, Hughes et al. *J Am Chem Soc* **106** 446 1984, Hughes et al. *J Org Chem* **52** 4745 1987, *Beilstein* **23** IV 2832.] See cinchonine [118-10-5].

**(11bR)-(–)- and (11bS)-(+)- 4,4-Dibutyl-4,5-dihydro-2,6-bis(3,4,5-trifluorophenyl)-3H-dinaphth[2,1-c:1',2'-e]azepinium bromide** [*R*- 887938-70-7; *S*- 851942-89-7] **M 748.6, m 223-228°**, **224-229°**,  $[\alpha]_{\text{D}}^{29} \pm 11.7^{\circ}$  (**c 1.01, CHCl<sub>3</sub>**). The precursor chiral 2,2'-bisbromomethyl -1,1'-binaphthyl (280mg, 0.4mmol), di-*n*-butylamine (140μl, 0.8mmol) and K<sub>2</sub>CO<sub>3</sub> (82mg, 0.6mmol) in MeCN (5ml) is refluxed with stirring for 10 hours. The mixture was poured into H<sub>2</sub>O, extracted with CH<sub>2</sub>Cl<sub>2</sub>, the extract is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated to dryness. The residue is purified by column chromatography on silical gel (eluting with MeOH/CH<sub>2</sub>Cl<sub>2</sub>, 1:20) to give the bromide salt (275mg, 92%) which may be recrystallised from MeOH/Et<sub>2</sub>O. The *S*-enantiomer has IR with  $\nu_{\text{max}}$  (film) at 2965, 2936, 2876, 1614, 1524, 1470, 1447, 1360, 1242, 1047, 923, 908, 870, 725 cm<sup>-1</sup> and <sup>1</sup>H NMR (400MHz, CDCl<sub>3</sub>) with  $\delta_{\text{H}}$  at 7.97-7.95 (4H, m, Ar-H), 7.55-7.51 (2H, m, Ar-H), 7.27-7.23 (8H, m, Ar-H), 4.99 (2H, d, *J* = 14.1Hz, Ar-CH<sub>2</sub>), 3.74 (2H, d, *J* = 14.1Hz, Ar-CH<sub>2</sub>), 3.32 (2H, t, *J* = 12.5Hz, N-CH<sub>2</sub>-CH<sub>2</sub>), 2.56 (2H, t, *J* = 12.3Hz, N-CH<sub>2</sub>-CH<sub>2</sub>), 1.06-0.97 (6H, m, CH<sub>2</sub>), 0.71 (6H, t, *J* = 6.9Hz, CH<sub>3</sub>), 0.23 (2H, bs, CH<sub>2</sub>); the <sup>13</sup>C NMR (100MHz, CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 151.03 (d, *J*<sub>C-F</sub> = 294.1Hz), 139.63 (ddd, *J*<sub>C-F</sub> = 254.1, 15.1, 15.1Hz), 138.30, 136.86, 134.63 (dd, *J*<sub>C-F</sub> = 12.7, 7.8Hz), 133.43, 131.13, 130.83, 128.30, 128.29, 127.68, 127.36, 123.32, 115.40-113.94 (m), 57.54, 57.34, 24.59, 19.32, 13.26. [Kitamura et al. *Angew Chem, Int Ed* **44** 1549 2005, Ooi et al. *J Org Chem* **68** 4576 2003.] These are potent phase transfer organocatalysts for asymmetric  $\alpha$ -alkylation of *N*-arylidene-glycine *tert*-butyl ester derivatives for the synthesis of chiral  $\alpha$ -substituted  $\alpha$ -amino acids at extremely low concentrations of catalyst [Ooi et al. *Tetrahedron Asymm* **17** 603 2006].

**Cetyltrimethylammonium bromide (cetrimonium bromide, CTAB)** [57-09-0] **M 364.5, m 227-235°(dec)**. Crystallise it from EtOH, EtOH/\*benzene or from wet acetone after extracting twice with petroleum ether. Shake it with anhydrous diethyl ether, filter and dissolve it in a little hot MeOH. After cooling in the refrigerator, the precipitate is filtered off at room temperature and re-dissolved in MeOH. Anhydrous ether is added and, after warming to obtain a clear solution, it is cooled and the crystalline material is collected. [Dearden & Wooley *J Phys Chem* **91** 2404 1987, Hakemi et al. *J Am Chem Soc* **91** 120 1987, *Beilstein* **4** IV 819.]

**Cetyltrimethylammonium chloride** [112-02-7] **M 320.0**. Crystallise the chloride from acetone/ether mixture, EtOH/ether, or from MeOH. [Moss et al. *J Am Chem Soc* **109** 4363 1987, *Beilstein* **4** IV 819.]

**Decyltrimethylammonium bromide** [2082-84-0] **M 280.3, m 239-242°**. Crystallise the salt from 50% (v/v) EtOH/Et<sub>2</sub>O, or from acetone and wash with ether. Dry it under vacuum at 60°. Also recrystallise it from EtOH and dry it over silica gel. [McDonnell & Kraus *J Am Chem Soc* **73** 2170 1952, Dearden & Wooley *J Phys Chem* **91** 2404 1987, *Beilstein* **4** IV 784.]

**Didodecyldimethylammonium bromide** [3282-73-3] **M 463.6, m 157-162°**. Recrystallise the salt from acetone, acetone/ether mixture, then from ethyl acetate, wash with ether and dry it in a vacuum oven at 60° [Chen et al. *J Phys Chem* **88** 1631 1984, Rupert et al. *J Am Chem Soc* **107** 2628 1985, Halpern et al. *J Am Chem Soc* **108** 3920 1986, Allen et al. *J Phys Chem* **91** 2320 1987]. [*Beilstein* **4** IV 801.]

**Diocetadecyldimethylammonium bromide** [3700-67-2] **M 630.9, m 161-163°**. Crystallise the bromide from acetone, then MeOH [Lukac *J Am Chem Soc* **106** 4387 1984]. Also purify it by chromatography on alumina by washing with \*C<sub>6</sub>H<sub>6</sub> and eluting with Me<sub>2</sub>CO, evaporating and crystallising from MeCN [Swain & Kreevoy *J Am Chem Soc* **77** 1126 1955]. [*Beilstein* **4** IV 829.]

***N,N*-Diocetadecyl methylamine (hydrogen ionophore III)** [4088-22-6] **M 536.0, m 40°, 44-46°, 48-49°, b 252-259°, pK<sub>Est</sub> ~10**. It can be distilled at high vacuum, but by dissolving in \*C<sub>6</sub>H<sub>6</sub>, filtering and evaporating, a waxy solid suitable for electrode use can be obtained. It recrystallises from Me<sub>2</sub>CO or MeCN. [Hoerr et al. *J Org Chem* **9** 201 1944, Wu & Yu *Talanta* **34** 577 1987, *Beilstein* **4** III 435.]

**Dodecyltrimethylammonium chloride** [112-00-5] **M 263.9, m 246<sup>o</sup>(dec)**. Dissolve the chloride in MeOH, treat with active charcoal, filter and dry it *in vacuo* [Waldenburg *J Phys Chem* **88** 1655 1984], or recrystallise it several times from 10% EtOH in acetone. It has also been repeatedly crystallised from EtOH/ether or MeOH. [Cella et al. *J Am Chem Soc* **74** 2062 1952, *Beilstein* **4** IV 79.]

**Hexyltrimethylammonium bromide** [2650-53-5] **M 224.3, m 186<sup>o</sup>**. Recrystallise it from acetone. It is extremely *hygroscopic*. [McDowell and Kraus *J Am Chem Soc* **73** 2170 1951, *Beilstein* **4** IV 710.]

**Octadecyl trimethylammonium bromide (stearyl trimethylammonium bromide)** [1120-02-1] **M 392.5, m ~250<sup>o</sup>dec, 230-240<sup>o</sup>(dec)**. Crystallise it from EtOH or H<sub>2</sub>O (solubility is 1 in 1000parts). It is very soluble in Me<sub>2</sub>CO. It is a bactericide. [Sheldon et al. *J Am Chem Soc* **68** 754 1946, Grieger & Kraus *J Am Chem Soc* **70** 3805, 38007, *Beilstein* **4** IV 827.]

**Tetra-*n*-amylammonium bromide (tetra-*n*-pentylammonium bromide)** [866-97-7] **M 378.5, m 100-101<sup>o</sup>**. Crystallise it from petroleum ether, \*benzene or acetone/ether mixtures and dry it *in vacuo* at 40-50<sup>o</sup> for 2 days. It is used in ion-paired chromatography. (Sagara et al. *J Chromatogr* **328** 289 1985). [*Beilstein* **4** IV 677.]

**Tetra-*n*-amylammonium iodide** [2498-20-6] **M 425.5, m 135-137<sup>o</sup>**. Crystallise the iodide from EtOH and dry it at 35<sup>o</sup> under a vacuum. It has also been purified by dissolving in acetone and precipitating by adding diethyl ether, and drying at 50<sup>o</sup> for 2 days. [*Beilstein* **4** IV 677.]

**Tetradecyltrimethylammonium bromide (myristyl trimethylammonium bromide)** [1119-97-7] **M 336.4, m 244-245<sup>o</sup>, 244-249<sup>o</sup>**. Crystallise the bromide from acetone or a mixture of Me<sub>2</sub>CO and >5% MeOH, or Me<sub>2</sub>CO/EtOH. Wash it with diethyl ether and dry it in a vacuum oven at 60<sup>o</sup>. It is a cationic detergent. Its solubility is 1g/5g H<sub>2</sub>O. [Dearden & Wooley *J Phys Chem* **91** 2404 1987, Shelton et al. *J Am Chem Soc* **68** 754 1946, *Beilstein* **4** III 419, **4** IV 813.]

**Tetra-*n*-heptylammonium iodide** [3535-83-9] **M 537.7, m 102-103<sup>o</sup>**. Crystallise the iodide from EtOH or aqueous EtOH. [Eriksen et al. *J Org Chem* **25** 849 1960, *Beilstein* **4** IV 736 for triheptylamine.]

**Tetra-*n*-hexylammonium bromide** [4328-13-6] **M 434.6, m 99-100<sup>o</sup>**. Wash the bromide with ether, and dry it in a vacuum at room temperature for 3 days.

**Tetra-*n*-hexylammonium chloride** [5922-92-9] **M 390.1**. Crystallise the chloride from EtOH.

**Tetra-*n*-hexylammonium iodide** [2138-24-1] **M 481.6, m 99-101<sup>o</sup>, 102-103<sup>o</sup>**. Wash the iodide with diethyl ether and dry it at room temperature *in vacuo* for 3 days. It is soluble in CH<sub>2</sub>Cl<sub>2</sub>. [Eriksen et al. *J Org Chem* **25** 849 1960, *Beilstein* **4** IV 711 for trihexylamine.]

**Tetrahexylammonium perchlorate** [4656-81-9] **M 454.1, m 104-106<sup>o</sup>**. Crystallise the salt from acetone and dry it *in vacuo* at 80<sup>o</sup> for 24 hours.

**Tetrapentylammonium bromide**. See tetra-*n*-amylammonium bromide above.

**Tetra-*n*-propylammonium bromide** [1941-30-6] **M 266.3, m >280<sup>o</sup>(dec)**. Crystallise it from ethyl acetate/EtOH (9:1), acetone or MeOH. Dry it at 110<sup>o</sup> under reduced pressure. [*Beilstein* **4** IV 471.]

**Tetra-*n*-propylammonium iodide** [631-40-3] **M 313.3, m >280<sup>o</sup>(dec)**. Purify the iodide by crystallising it from EtOH, EtOH diethyl ether (1:1), EtOH water or aqueous acetone. Dry it at 50<sup>o</sup> under a vacuum and store it over P<sub>2</sub>O<sub>5</sub> in a vacuum desiccator. Store it away from light. [*Beilstein* **4** IV 472.]

**Tri-*n*-butylhexadecylphosphonium bromide (TBHDTB, hexadecyltributylphosphonium bromide)** [14937-45-2] **M 507.7, m 54<sup>o</sup>, 56-58<sup>o</sup>**. It is made by heating 1-bromohexadecane (1 mol, 112-82-3) and tri-*n*-butylphosphine (1 mol, 998-40-3) at 65<sup>o</sup> for 3 days. The mixture solidifies on cooling, and the solid is recrystal-

lised from hexane and dried *in vacuo* to give 68% of the phosphonium salt. [Starks *J Am Chem Soc* **93** 195 1971, *cf* for preparing fluorides see Landini et al. *Synthesis* 37 and 428 1974, for preparing disulfides see Landini & Rolla *Synthesis* 565 1974.]

**Tri-*n*-dodecylamine (Hydrogen ionophore I)** [102-87-4] **M 522.0, m 15.7°, b 220-228°/0.03mm, d<sub>4</sub><sup>20</sup> 0.833, n<sub>D</sub><sup>20</sup> 1.4577, pK<sub>Est</sub> ~11.0.** Distil tridodecylamine under high vacuum and N<sub>2</sub>, and store it in the absence of CO<sub>2</sub>. It can be recrystallised from 95%EtOH/\*C<sub>6</sub>H<sub>6</sub> at low temperature under vacuum. The *hydrochloride* has **m 78-79°**. [Ra et al. *J Org Chem* **9** 259 1944, *Beilstein* **4** III 413, **4** IV 801.]

**Tri-*n*-dodecylammonium nitrate** [2305-34-2] **M 585.0.** Crystallise the salt from *n*-hexane/acetone (95:5) and keep it in a desiccator over P<sub>2</sub>O<sub>5</sub> under vacuum. [*Beilstein* **4** IV 801 for tridodecylamine.]

**Tri-*n*-dodecylammonium perchlorate** [5838-82-4] **M 622.4.** Recrystallise the salt from *n*-hexane or acetone and keep it in a desiccator over P<sub>2</sub>O<sub>5</sub>. (Potentially explosive.)

**Tri-*n*-octylammonium chloride** [1188-95-0] **M 384.2, m 78-79°, pK<sup>25</sup> 8.35 (in 70% aqueous EtOH).** Crystallise it from Et<sub>2</sub>O, then *n*-hexane (see above). [Burrows et al. *J Chem Soc* 200 1947, *Beilstein* **4** H 196.]

**Tri-*n*-octylammonium perchlorate** [2861-99-6] **M 454.2, m >300°(dec).** Crystallise the perchlorate from *n*-hexane. (Possibly explosive.) [*Beilstein* **4** IV 754.]

## IMIDAZOLIUM IONIC LIQUID CRYSTAL CATALYSTS

A general discussion on ionic liquid crystals has already been made (see Chapter 3), and this section includes some commercially available liquid crystal 1,3-dialkylimidazolium salts that participate in the catalytic process, i.e. by coordination with the metal component of catalysts, as well as assisting in other ways such as acting as a solvent, affecting the dielectric, and in phase transference of reagents. Depending on the *N*-alkyl substituents and the counter anion the salts may be more, or less, soluble in water or organic solvents. They are non-volatile (i.e. environmentally friendly), stable to water and air and many may be heated to temperatures as high as 300°, making it possible to distil off some products from them. They can be recovered for re-use. [See P. Wasserscheid and T. Welton, *Ionic Liquids in Synthesis* (2 volume set), Wiley VCH, Weinheim, 2008. ISBN-10: 3527312390, ISBN-13: 978-3527312399.]

*Preparation:* Alkylation of *N*-1 of imidazole can be carried out using alkyl halide and alkoxide as for 1-ethylimidazole [1072-62-4], and the second alkylation onto *N*-3 is performed by further reaction with the desired alkylbromide (slightly less than a molar equivalent) in a solvent, e.g. 1,1,1-trichloroethane, at reflux for 2 hours. The molten bromide salt is separated from the solvent, washed with trichloroethane, dried on a Rotavap at 70°/0.1mm to give the bromide salt as a white solid or a liquid in high yield (~90%) and high state of purity if the starting reagents are pure. Check purity by <sup>1</sup>H NMR and elemental analysis. The imidazolium bromide salts are most useful as they can be converted into other salts, e.g. AcO<sup>-</sup> with AgOAc, TfO<sup>-</sup> with AgOTf, bis(trifluoromethylsulfonyl)imide [NTf<sub>2</sub><sup>-</sup>] with NTf<sub>2</sub>Li, nonafluorobutanesulfonate [NfO<sup>-</sup>] with NfOK, etc. Generally the imidazolium bromide and the metal salt of the required anion are dissolved in the least volume of H<sub>2</sub>O at 70°, stirred for 1 hour (if silver halide precipitates it is filtered off), and the imidazolium salt can be extracted into an organic solvent, e.g. CH<sub>2</sub>Cl<sub>2</sub> or 1,1,1-trichloroethane, and the extract is then evaporated *in vacuo* and dried at 50°/0.1mm for 2 hours or until there is no further loss in weight. *Alternatively*, where possible, the 1-alkylimidazole is alkylated with e.g. alkyl toluene-*p*-sulfonate ester to provide the 1,3-dialkylimidazolium toluene-*p*-sulfonate directly. [Bonhôte et al. *Inorg Chem* **35** 1168 1996.]

The following liquid crystal salts can be readily prepared in small or large quantities:

**1-Butyl-3-methylimidazolium Salts (BMIM<sup>+</sup> X<sup>-</sup>,** see Park & Kazlauskas *J Org Chem* **66** 8395 2001 for improved preparation, purification and use at ambient temperature of some of the following ionic liquids in lipase-catalysed enantio- and regioselective acylation reactions:

**BMIM<sup>+</sup> AcO<sup>-</sup>** [284049-75-8] **M 198.6; BMIM<sup>+</sup> Ntf<sub>2</sub><sup>-</sup>** [174899-83-3] **M 419.4, n<sub>D</sub><sup>20</sup> 1.428,** [good synthetic activity of soluble *Candida antarctica* lipase B in this ionic liquid was obtained towards the enantioselective and operational stability for butyl butyrate synthesis and kinetic resolution of 1-phenylethanol in supercritical CO<sub>2</sub>,

Lozano et al. *JCS, Chem Commun* 692 2002]; **BMIM<sup>+</sup> Br<sup>-</sup>** [85100-77-2] **M 219.1**, [used in Heck reactions with Pd(OAc)<sub>2</sub> where Pd-Im bond is formed, Xu et al. *Organometallics* **19** 1123 2000]; **BMIM<sup>+</sup> Cl<sup>-</sup>** [79917-90-1] **M 174.7, m ~70°**, [used with AlCl<sub>3</sub> in Friedel-Crafts alkylation reactions, Chauvin et al. *J Mol Catal* **92** 155 1994]; **BMIM<sup>+</sup> (n-BuO)<sub>2</sub>P(O)O<sup>-</sup>** [663199-28-8] **M 348.4, n<sub>D</sub><sup>20</sup> 1.472**; **BMIM<sup>+</sup> (NC)<sub>2</sub>N<sup>-</sup>** [448245-52-1] **M 205.3**; **BMIM<sup>+</sup> SbF<sub>6</sub><sup>-</sup>** [174645-81-9] **M 375.0**; and **BMIM<sup>+</sup> PF<sub>6</sub><sup>-</sup>** [174501-64-5] **M 284.2, n<sub>D</sub><sup>20</sup> 1.411**, [used in allylation of RCHO with (Allyl)Sn, Gordon & McCluskey *JCS, Chem Commun* 1431 1999; in coupling of aryl halides by (Ph<sub>3</sub>)<sub>n</sub>Ni(0), Howarth et al. *Tetrahedron Lett* **41** 10319 2000; and catalyse the addition of CN (from TMSCN) to arylimines (from ArCHO + RNH<sub>2</sub>) to form α-aminonitriles, Yadav et al. *New J Chem* **27** 462 2003]; **BMIM<sup>+</sup> HCO<sub>3</sub><sup>-</sup>** [366491-15-8] **M 200.2, as 50% in 2:3 MeOH:H<sub>2</sub>O**; **BMIM<sup>+</sup> HSO<sub>4</sub><sup>-</sup>** (as **BASIONIC<sup>®</sup> AC 28**) [262297-13-2] **M 236.3**; **BMIM<sup>+</sup> MeSO<sub>3</sub><sup>-</sup>** (as **BASIONIC<sup>®</sup> ST 78**) [342789-81-5] **M 234.3**; and **BMIM<sup>+</sup> MeSO<sub>4</sub><sup>-</sup>** [401788-98-5] **M 250.3, n<sub>D</sub><sup>20</sup> 1.478**, [catalyse the addition of CN (from TMSCN) to arylimines (from ArCHO + RNH<sub>2</sub>) to form α-aminonitriles, Yadav et al. *New J Chem* **27** 462 2003]; **BMIM<sup>+</sup> BF<sub>4</sub><sup>-</sup>** [174501-65-6] **M 226.0**, {specifically catalyses the Biginelli reaction (formation of 3,4-dihydropyrimidin-2(1H)-ones from aldehydes + urea + MeCOCH<sub>2</sub>COR), Peng & Deng *Tetrahedron Lett* **42** 5917 2001; **BMIM<sup>+</sup> PF<sub>6</sub><sup>-</sup>** [174501-64-5] **M 284.2**, [catalyses the addition of CN (from TMSCN) to arylimines (from ArCHO + RNH<sub>2</sub>) to form α-aminonitriles, Yadav et al. *New J Chem* **27** 462 2003; assists in the bi-phasic hydrogenation of arenes at room temperature using a ruthenium cluster catalyst which coordinates with it, Dyson et al. *JCS, Chem Commun* 25 1999; and the molten salt catalyses the asymmetric hydrogenation of 2-arylacrylic acids by immobilised Ru-BINAP complex, Montiero et al. *Tetrahedron Asymmetry* **8** 177 1997; and is used in Palladium [Pd(Ph<sub>3</sub>)<sub>4</sub>] catalysed Suzuki cross-coupling (aryl halides and arylboronic acids) at ambient temperature, Mathews et al. *JCS, Chem Commun* 1249 2000; and it specifically catalyses the Biginelli reaction (formation of 3,4-dihydropyrimidin-2(1H)-ones from aldehydes + urea + MeCOCH<sub>2</sub>COR), Peng & Deng *Tetrahedron Lett* **42** 5917 2001; and **BMIM<sup>+</sup> CF<sub>3</sub>SO<sub>3</sub><sup>-</sup>** [174899-66-2] **M 288.3, n<sub>D</sub><sup>20</sup> 1.434**; **BMIM<sup>+</sup> octyOSO<sub>3</sub><sup>-</sup>** [445473-58-5] **M 348.5**, [this ionic liquid increases the yield and enzyme stability of β-galactosidase in enzyme-catalysed syntheses, Kaftzik et al. *Org Process Res Dev* **6** 553 2002].

**1-Ethyl-3-methylimidazolium Salts (EMIM<sup>+</sup> X<sup>-</sup>**, see Park & Kazlauskas *J Org Chem* **66** 8395 2001 for improved preparation and use of ambient temperature of some of the following ionic liquids in lipase-catalysed enantio- and regioselective acylation reactions):

**EMIM<sup>+</sup> Br<sup>-</sup>** [65039-08-9] **M 191.1**, [it is highly conductive at ~25° [Bonhôte et al. *Inorg Chem* **35** 1168 1996]; **EMIM<sup>+</sup> Cl<sup>-</sup>** [65039-09-0] **M 146.6, m 77-79°**; **EMIM<sup>+</sup> (n-BuO)<sub>2</sub>P(O)O<sup>-</sup>** [869858-84-4] **M 320.4, n<sub>D</sub><sup>20</sup> 1.469**; **EMIM<sup>+</sup> AcO<sup>-</sup>** [143314-17-4] **M 170.2, n<sub>D</sub><sup>20</sup> 1.502**; **EMIM<sup>+</sup> (CF<sub>3</sub>CF<sub>2</sub>-SO<sub>2</sub>)<sub>2</sub>N<sup>-</sup>** [216299-76-2] **M 491.3, m ≤1°**, [very stable fluorinated ionic liquids that are extremely hydrophobic]; **EMIM<sup>+</sup> Ntf<sub>2</sub><sup>-</sup>** [174899-82-2] **M 391.3, m ≥15°**, [it is a useful medium for the enantioselective cyclopropanation of styrene with ethyldiazoacetate, promoted by two different Cu-bis(oxazoline) complexes and its recovery, Fraile et al. *Tetrahedron Asymmetry* **12** 1891 2001; and good synthetic activity of soluble *Candida antarctica* lipase B in this ionic liquid was obtained towards the enantioselective and operational stability for butyl butyrate synthesis and kinetic resolution of 1-phenyl ethanol in supercritical CO<sub>2</sub>, Lozano et al. *JCS, Chem Commun* 692 2002]; **EMIM<sup>+</sup> (CN)<sub>2</sub>N<sup>-</sup>** [370865-89-7] **M 177.2**, [its conductivity is ~26,000 μS cm, and its electrochemical window is -2.4 to +3.3 V, cf Bonhôte et al. *Inorg Chem* **35** 1168 1996]; **EMIM<sup>+</sup> PF<sub>6</sub><sup>-</sup>** [155371-19-0] **M 256.1, m 58-62°**, [it is prepared by mixing EMIM<sup>+</sup> Cl<sup>-</sup> (29.3g, 200mmol) and 60% aqueous HPF<sub>6</sub> (9g, 200mmol) in H<sub>2</sub>O (300ml), the resulting mixture of white solid and liquid are cooled in an ice bath for 2 hours and the EMIM<sup>+</sup> PF<sub>6</sub><sup>-</sup> (31.8g 62%) is dried *in vacuo*. Recrystallisation from MeOH provides crystals for X-ray structural analysis. [Fuller et al. *JCS, Chem Commun* 299 1994]; and **EMIM<sup>+</sup> BF<sub>4</sub><sup>-</sup>** [143314-16-3] **M 198.0, m 15°, b >350°, d<sub>4</sub><sup>25</sup> 1.294, n<sub>D</sub><sup>20</sup> 1.413**, [its conductivity is ~11,500 μS cm, and its electrochemical window is -2.2 to +3.5 V, cf Bonhôte et al. *Inorg Chem* **35** 1168 1996]; the salt is prepared by stirring Ag<sub>2</sub>O (23.2g, 100mmol) with 48% aqueous HBF<sub>4</sub> (36.9g, 200mmol) in H<sub>2</sub>O (300ml) until the Ag<sub>2</sub>O has reacted completely to give a clear solution, then EMIM<sup>+</sup> Cl<sup>-</sup> (29.2g, 200mmol) dissolved in H<sub>2</sub>O is added and the mixture is stirred for 2 hours, the AgCl is filtered off, the filtrate is evaporated *in vacuo* and the colourless residue is dried in a vacuum oven at 60° to give the BF<sub>4</sub><sup>-</sup> salt (33.6g, 85%). [Fuller et al. *JCS, Chem Commun* 299 1994].

## CHIRAL AUXILIARIES

These compounds are involved in reactions, which direct stereospecificity and are decomposed, e.g. by

hydro-lysis, reduction, to generate the desired chiral products and the remains of the original auxiliary which can be recycled. Unlike the case of catalysis, the auxiliary is involved stoichiometrically and is not recycled during the reaction. It is one of the products of the reaction and can, *via* other reactions if necessary, be used to regenerate the original auxiliary. These have been used for decades and are of necessity chiral molecules themselves. The auxiliaries have to be linked to the molecule that has a reactive group which can be operated on, e.g. carbonyl, to generate an asymmetric centre, e.g. a chiral alcohol. This section includes more recently available auxiliaries, although many useful ones such as cinchonine, cinchonidine, borneols, camphors, menthol derivatives, substituted sugars, substituted chiral heterocycles (e.g. pyrrolidines), sterols etc. will be found scattered in Chapters 4, 5 and 7. [see Evans *Aldrichimica Acta* **15** 23 1982, Ager et al. *Chem Rev* **96** 835 1996, Ager et al. *Aldrichimica Acta* **30** 3 1997, Mukaiyama *Aldrichimica Acta* **29** 59 1996, K. Rück-Braun and H. Kunz *Chiral Auxiliaries in Cycloaddition Reactions* Wiley-VCH, Weinheim 1999, ISBN 3-527-29386-8.]

***R*-(–)- and *S*-(+)- Acetyl-cyclopentadienyl-ironcarbonyl triphenylphosphine complex [*R*- and *S*- ( $\eta^5$ -C<sub>5</sub>H<sub>5</sub>)Fe(CO)(CH<sub>3</sub>CO) (PPh<sub>3</sub>)]** [*R* 36548-61-5; *S* 36548-60-4; *RS* 12101-02-9] **M 454.3, m 142°**, (+), **140°** (–), **145°** ( $\pm$ ), ***R*-[ $\alpha$ ]<sub>436</sub><sup>27</sup> –288°, [ $\alpha$ ]<sub>546</sub><sup>22</sup> –288° (c 0.045, C<sub>6</sub>H<sub>6</sub>)**, ***S*- [ $\alpha$ ]<sub>546</sub><sup>22</sup> +288° (c 0.045, C<sub>6</sub>H<sub>6</sub>)**. This racemic “chiral” auxiliary can be made on a large scale from [ $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)<sub>2</sub>]<sub>2</sub> [38117-54-3] which is cleaved with Na/Hg to give the anion [ $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)<sub>2</sub>]<sup>–</sup> (Na<sup>+</sup>) that is methylated to  $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)COMe (m **78–82°**, caramel coloured waxy crystals from sublimation *in vacuo* onto an ice-cooled glass finger) with MeI [Piper & Wilkinson *J Inorg Nucl Chem* **3** 104 1956, Aktogu et al. *J Organomet Chem* **262** 49 1984, King *Organometallic Synthesis Vol 1*, Academic Press, NY, p 145 1965]. This methyl complex provides the racemic title compound when treated with PPh<sub>3</sub>. Thus (C<sub>5</sub>H<sub>5</sub>)Fe(CO)<sub>2</sub>Me (1.0g, 5mmol) and PPh<sub>3</sub> (1.3g, 5mmol) in redistilled THF (10ml) are refluxed (65°) under N<sub>2</sub> until the Fe-Me band (IR: strong C-H deformation band at 1170 cm<sup>–1</sup> in CS<sub>2</sub>) disappears (~48 hours). The solution is filtered, the solvent is evaporated (to ~20mm), the residue is dissolved in pentane (10ml), passed through an Al<sub>2</sub>O<sub>3</sub> column (5 x 20cm), and only one band (yellow orange to orange) is eluted with pentane which, on evaporation (at ~20mm), provides the analytically pure orange ( $\pm$ )- $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)(CH<sub>3</sub>CO)(PPh<sub>3</sub>), m 145° in ~98% yield. This reaction is solvent dependent, i.e. no reaction occurs in boiling hexane (68°), and is only 50% complete in boiling Et<sub>2</sub>O (34°) after 48 hours. The solid is stable in air, is soluble in organic solvents, e.g. pentane, hexane, Et<sub>2</sub>O, THF, CH<sub>2</sub>Cl<sub>2</sub>, CHCl<sub>3</sub>, and \*C<sub>6</sub>H<sub>6</sub>, but is insoluble in MeOH or H<sub>2</sub>O. Solutions in CHCl<sub>3</sub>, and \*C<sub>6</sub>H<sub>6</sub> decompose rapidly in air to produce brown intractable solids. Hence these solutions should be prepared and used under N<sub>2</sub> or argon. Its IR (CHCl<sub>3</sub>) has bands with  $\nu_{\max}$  at 1598 (s, MeC=O) and 1920 (vs, br, Fe carbonyl) cm<sup>–1</sup>; and the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, external TMS) has  $\delta_{\text{H}}$  at 7.59 (m, Ph, 15H), 4.69 (s, C<sub>5</sub>H<sub>5</sub>, 5H) and 2.52 (s, COMe, 3H). [Bibler & Wojcicki *Inorg Chem* **5** 889 1966, Butler et al. *Inorg Chem* **6** 2074 1967.]

The versatility of this acetyl auxiliary has prompted its **optical resolution** into the pure enantiomers which proved to be very good chiral auxiliaries for preparing a variety of optically active molecules where high stereo control is achieved. Two independent resolutions were achieved, both involving 1*R*,2*S*,5*R*-(–)-2-isopropyl-5-methylcyclohexan-1-ol (*R*-(*l*-menthol). The first is from the reaction of sodium *R*-menthylate (NaOC<sub>10</sub>H<sub>19</sub>, *R*-menthyl refers to the radical produced by loss of the hydroxyl group) and ( $\pm$ )-( $\eta^5$ -C<sub>5</sub>H<sub>5</sub>)Fe(CO)<sub>2</sub>(PPh<sub>3</sub>)<sup>+</sup>PF<sub>6</sub><sup>–</sup> (see below) recovered from hydrolysis, decomplexation, of derivatives of the racemic title compound Aktogu et al. [*J Organomet Chem* **262** 49 1984] to give a diastereoisomeric mixture of (+)-(C<sub>5</sub>H<sub>5</sub>)Fe(CO)(PPh<sub>3</sub>)-(–)COOC<sub>10</sub>H<sub>19</sub> {[ $\alpha$ ]<sub>589</sub><sup>20</sup> +30°, [ $\alpha$ ]<sub>579</sub><sup>20</sup> +35°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +70°, [ $\alpha$ ]<sub>436</sub><sup>20</sup> 1450° (in 10<sup>–3</sup>M \*C<sub>6</sub>H<sub>6</sub>)}, and (–)-(C<sub>5</sub>H<sub>5</sub>)Fe(CO)(PPh<sub>3</sub>)-(–)COOC<sub>10</sub>H<sub>19</sub> {[ $\alpha$ ]<sub>589</sub><sup>20</sup> –75°, [ $\alpha$ ]<sub>579</sub><sup>20</sup> –80°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> –120°, [ $\alpha$ ]<sub>436</sub><sup>20</sup> +1550° (in 10<sup>–3</sup>M \*C<sub>6</sub>H<sub>6</sub>)}, (together with NaPF<sub>6</sub>) which were separated by recrystallisation from pentane, with the latter being more soluble. [Brunner & Schmidt *J Organomet Chem* **21** P53 1970.] Then the ester with [ $\alpha$ ]<sub>546</sub><sup>20</sup> –120° (560mg, 0.93mmol) in THF (20ml) at –30°, is treated dropwise with MeLi (1ml of 1.5M Et<sub>2</sub>O solution), followed by stirring at –30°/1 hour, then at ~25°/1 hour. After quenching the reaction and evaporating, the brown residue is extracted into \*C<sub>6</sub>H<sub>6</sub>, and purified through a column of Al<sub>2</sub>O<sub>3</sub>/3% H<sub>2</sub>O and eluted with \*C<sub>6</sub>H<sub>6</sub>. The greenish zone gives a menthol-free yellow solid, which on sublimation provides analytically pure orange (+)- $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)(CH<sub>3</sub>CO)(PPh<sub>3</sub>) (47mg, 11% yield), m 142°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> –228° (in 10<sup>–3</sup>M \*C<sub>6</sub>H<sub>6</sub>), found M 465 (osmometry in \*C<sub>6</sub>H<sub>6</sub>). Similarly the menthyl ester with [ $\alpha$ ]<sub>546</sub><sup>20</sup> +70° gives pure (–)- $\eta^5$ -(C<sub>5</sub>H<sub>5</sub>)Fe(CO)(CH<sub>3</sub>CO)(PPh<sub>3</sub>) (80mg, 30% yield), m 140°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +227° (in 10<sup>–3</sup>M \*C<sub>6</sub>H<sub>6</sub>), found M 464 (osmometry in \*C<sub>6</sub>H<sub>6</sub>). From the CD and ORD spectra it was concluded that these reactions occurred with **inversion** of configuration at the tetrahedral iron centre [Brunner & Schmidt *J Organomet Chem* **36** C18 1972].

In the second resolution *R*-(-)-chloromethylmenthyl ether (*R*- is the configuration at 1-C-OH of menthol, see [26127-08-2]) was reacted with the lithium salt of  $(\pm)\text{-}\eta^5\text{-}(C_5H_5)Fe(CO)(CH_3CO)(PPh_3)$  (generated with *n*-BuLi) to give the diastereomeric ethers  $(-)\text{-}\eta^5\text{-}(C_5H_5)Fe(CO)(PPh_3)(COCH_2CH_2O(-)-C_{10}H_{19})$   $\{[\alpha]_D^{20} +65^\circ$  (c 0.4, \*C<sub>6</sub>H<sub>6</sub>) $\}$ , and  $(+)\text{-}\eta^5\text{-}(C_5H_5)Fe(CO)(PPh_3)(COCH_2CH_2O(-)-C_{10}H_{19})$   $\{[\alpha]_D^{20} -150^\circ$  (c 0.4, \*C<sub>6</sub>H<sub>6</sub>) $\}$  which are separable by chromatography and distinguishable by <sup>1</sup>H NMR (300MHz). Crystals of the latter diastereoisomer were subjected to x-ray crystallographic analysis which revealed that the absolute configuration at the tetrahedral Fe centre was *R* by virtue that the absolute configuration of the *R*-menthyl moiety had been established. Since the original formation of these ethers occurs with alteration in the configuration at the Fe centre, then the desired *S*(+)- and *R*(-)- configurations of  $(\pm)\text{-}\eta^5\text{-}(C_5H_5)Fe(CO)(CH_3CO)(PPh_3)$  are established. It was also shown that Brunner & Schmidt's reactions of the *R*-menthyl esters with *n*-BuLi to provide the title enantiomers occurred with *complete* inversion of configuration (as determined by 500MHz <sup>1</sup>H NMR spectroscopy). The enantiomers can be discriminated in solution (9mg in 700μl of CDCl<sub>3</sub>) containing the chiral shift reagent Eu(tfc) {tris[3-(trifluoromethylhydroxymethylene)-(+)-camphorato] europium (III), see [34834-11-0]} (0.48mg in 8μl) where a clear separation of the methyl singlets in the 300MHz <sup>1</sup>H NMR spectra occurs: the *S*(+)-enantiomer has δ at 2.66 and the *R*(-)-enantiomer has δ at 2.58 (from TMS). [Davies et al. *JCS Chem Commun* 607 1986.]

The chiral auxiliary properties of the complexes  $\eta^5\text{-}C_5H_5Fe(CO)(CH_3CO)(PPh_3)$  are displayed by deprotonation with *n*-BuLi (isoPr<sub>2</sub>NLi in THF has also been used) to give the blood-red enolate which undergoes a variety of reactions and sequences of reactions (homochiral synthesis) that are highly stereoselective. These include alkylation, homochiral succinylation, asymmetric aldol reactions, formation of α,β-unsaturated iron acyls and homochiral dienolates, asymmetric Michael addition reaction, asymmetric synthesis of β-amino acids and β-lactams, asymmetric synthesis of cyclopropanecarboxylic acids, asymmetric synthesis of chiral sulfoxides, and chirality recognition for homochiral synthesis. [Davies *Aldrichimica Acta* 23 31 1990.] Decomplexation at the end of the reactions can be achieved by oxidative cleavage (Br/MeOH, H<sub>2</sub>O or *n*-BuNH<sub>2</sub>, NBS/EtOH) to provide the acid, amide or ester, also by alcoholysis (to provide different esters) from which the products can be isolated. From the aqueous solution  $\eta^5\text{-}C_5H_5Fe(CO)_2(PPh_3)^+PF_6^-$  (evaporated to 20ml, from 1mmol of complex product) can be recovered by addition of NH<sub>4</sub>PF<sub>6</sub> (0.6g, 3.7mmol), stirring for 1 hour, evaporating to dryness, and the residue is extracted with CH<sub>2</sub>Cl<sub>2</sub> (3 x 10ml). The combined extracts are evaporated to 5ml and Et<sub>2</sub>O is added to give pale yellow crystals of the complex salt (0.38g, 65%, m 186°). It can also be recrystallised from Me<sub>2</sub>CO/Et<sub>2</sub>O or Me<sub>2</sub>CO/hexane. It has IR (Nujol) with ν<sub>max</sub> at 2055 and 2010 (terminal CO, in CHCl<sub>3</sub>) and 1435s, 1090s, 995w, 880m, 850-820vs, br, 740m, 700m, 685m, 605m, 575m, 555m, 525m, 500w cm<sup>-1</sup>; and the <sup>1</sup>H NMR (60MHz, Me<sub>2</sub>CO, TMS) has δ<sub>H</sub> at 5.62 (d, J<sub>P,H</sub> = 1.5Hz, C<sub>5</sub>H<sub>5</sub>) and 7.68 (m, Ph, 15H). [Aktogu et al. *J Organomet Chem* 262 49 1984, Treichel et al. *Inorg Chem* 5 1177 1966.]

***1R,2S,5R*(-)-Chloromethylmenthyl ether** [*R*(-)-chloromethylmenthyl ether] [26127-08-2] **M 204.7, b 62°/0.1mm, 160-162°/13-16mm, d 0.9821, d<sup>25</sup> 0.994, n<sub>D</sub><sup>20</sup> 1.467, n<sub>D</sub><sup>27</sup> 1.465, [α]<sub>D</sub><sup>24</sup> -177.0°** (c 1, CH<sub>2</sub>Cl<sub>2</sub>), and ***1S,2R,5S*(+)-chloromethylmenthyl ether** [*S*(+)-chloromethylmenthyl ether] [103128-76-3] **M 204.7, b 72°/0.4mm, d<sup>21</sup> 0.994, n<sub>D</sub><sup>20</sup> 1.467, [α]<sub>D</sub><sup>24</sup> +181.0°** (c 2, CH<sub>2</sub>Cl<sub>2</sub>). The chloro methyl derivatives of *R*(-)- and *S*(+)- menthol respectively are useful chiral auxiliary resolving agents (see previous entry) as they can be readily removed by mild hydrolysis and the recovered menthol can be recycled. The *R*(-)-enantiomer is prepared by melting *l*-menthol (100g, 0.64moles) on a water bath and stirring vigorously with 40% w/w aqueous formalin (50g, 0.67moles) while HCl gas is bubbled through. The mixture warms up at first and has to be cooled at 0° until no more gas dissolves. The clear mixture separates into two layers, the upper layer is removed and the lower layer is dried over Na<sub>2</sub>SO<sub>4</sub>. This is filtered and fractionated under reduced pressure from some unreacted menthol, formaldehyde trimer and HCl to give the *chloromethyl ether* (120-130g, 91-99%) as a slightly refracting oil which distils steadily at 160-162°/16mm (and 62°/0.1mm), but it decomposes at ca 230°/atm. The pure ether has [α]<sub>D</sub><sup>21</sup> -172.75° (c 6.78 CHCl<sub>3</sub>). On heating in EtOH/charcoal it forms the *dimethylmethylal* C<sub>10</sub>H<sub>19</sub>OCH<sub>2</sub>OC<sub>10</sub>H<sub>19</sub> which crystallises from aqueous EtOH in greasy looking scales or from Et<sub>2</sub>O in colourless needles m 57° (b 337°/atm), with [α]<sub>D</sub><sup>24</sup> -77.94° (c 0.8 EtOH). The *chloromethylmenthyl ethers* hydrolyse in H<sub>2</sub>O to menthol, HCHO and HCl so they should be stored in an inert atmosphere in the cold [Wedekind *Chem Ber* 34 813 1901, Deutsches Reichspatent No. 119008, D.R.P. 189331 *Chem Zentralblatt* 1 184 1908]. [Beilstein 6 H 32, 6 I 21.]

***4R*(+)-4-Isopropyl-5,5-dimethyl-2-oxazolidinone** (*4R*(+)-4-isopropyl-5,5-dimethyl-1,3-oxazolidin-2-one) [223906-38-5] and ***4S*(-)-4-isopropyl-5,5-dimethyl-2-oxazolidinone** (*4S*(-)-4-isopropyl-5,5-dimethyl-1,3-

**oxazolidin-2-one** [168297-86-7] **M 157.2, m 86-87°**,  $[\alpha]_D^{20} \pm 47^\circ$  (c 1, H<sub>2</sub>O). The starting chiral ethanolamine is prepared from the methyl esters of chiral  $\alpha$ -aminoacids (*R*- or *S*- valine in this case) by reaction with MeMgI/Et<sub>2</sub>O to give chiral 2,2-dimethylvalinol which is converted into the desired chiral 2-oxazolidinones with (EtO)<sub>2</sub>CO/K<sub>2</sub>CO<sub>3</sub> (as in the following entry), or by reaction with CCl<sub>3</sub>COCl (in pyridine) or carbonyl-diimidazole (in CH<sub>2</sub>Cl<sub>2</sub>) as carbonyl equivalents. Similarly by using the esters of chiral alanine, norleucine or  $\alpha$ -phenylglycine the respective optically active 5,5-dimethyl-2-oxazolidinones where the 4-isopropyl group is replaced by methyl, *n*-butyl and phenyl groups respectively can be prepared. These auxiliaries have been named "second series Quat auxiliaries", the first series being substituted 3,3-dimethyl-2-pyrrolidinones [Davies et al. *Tetrahedron Lett* **35** 2369, 2373 1994]. The present 2-oxazolidinones can be *N*-acylated, (e.g. with BuLi, then RCH<sub>2</sub>COCl or MeCH=CHCOCl), and the *N*-acyl moieties can be the targets for highly stereoselective enolate alkylation and conjugate addition reactions. The products can be hydrolysed (e.g. with LiOH, THF/H<sub>2</sub>O 3:1, 0° at ~25°) to provide the respective chiral acids and regenerated oxazolidin-2-one. The *gem*-dimethyl groups enhance the face-stereoselective shielding of the attached *N*-acyl moiety leading to very high diastereomeric excess in the products. [Davies & Sanganeer *Tetrahedron Asymm* **6** 671 1995, cf: review by Mukaiyama *Aldrichimica Acta* **29** 59 1996.]

**4R-(+)-4-Isopropyl-2-oxazolidinone (4R-(+)-4-isopropyl-1,3-oxazolidin-2-one)** [95530-58-8] and **4S-(-)-4-isopropyl-2-oxazolidinone (4S-(-)-4-isopropyl-1,3-oxazolidin-2-one)** [17016-83-0] **M 129.2, m 70-71.5°, 70-72°, 71-72°, R-  $[\alpha]_D^{20} +18^\circ$ , S-  $[\alpha]_D^{20} -18^\circ$ , (c 6, EtOH)**. These compounds are Evans' type of chiral auxiliaries. The *S*-enantiomer was prepared by stirring 1mol of *S*-valinol, 1.1mol of diethyl carbonate and 1mol of anhydrous K<sub>2</sub>CO<sub>3</sub> at 125-126° (internal temperature) until 2.0mols of EtOH had distilled off (4-6 hours). The cooled mixture (to 20°) is dissolved in Et<sub>2</sub>O, filtered through a pad of Celite to remove the K<sub>2</sub>CO<sub>3</sub>, evaporated to a small volume and cooled slowly to 0° when the oxazolinone crystallises as white needles (m 69-70°, 85-95% yield). It is soluble in CH<sub>2</sub>Cl<sub>2</sub> and recrystallises from hexanes/EtOAc (4:1, v/v) by allowing it to stand at 6° overnight. On TLC (0.25mm silica gel 60-F plates) it has R<sub>F</sub> 0.19 (hexanes/EtOAc 6:4,v/v). It has  $[\alpha]_{589} -16.6^\circ$ ,  $[\alpha]_{577} -17.3^\circ$ ,  $[\alpha]_{546} -20.2^\circ$ ,  $[\alpha]_{435} -37.3^\circ$ ,  $[\alpha]_{365} -63.7^\circ$  (c 5.81, EtOH); the IR (CH<sub>2</sub>Cl<sub>2</sub>) has  $\nu_{\max}$  at 1240, 1400, 1760, 2980, 3060, 3240-3340, 3480 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (90MHz, CDCl<sub>3</sub>) has  $\delta$  at 6.7 (br s, 1H, NH), 4.42 (t, *J* = 8.6Hz, 1H, C<sub>5</sub>-H), 4.07 (d of d, *J* = 8.5, 6.5Hz, 1H, C<sub>5</sub>-H), 3.58 (d of t, *J* = 8.6, 6.5Hz, 1H, C<sub>4</sub>-H), 1.9-1.6 (m, 1H, C<sub>4</sub>-H), 0.95 (overlapping d's, *J* = 6.0Hz, 6H, CH(CH<sub>3</sub>)<sub>2</sub>). [Evans et al. *J Org Chem* **50** 1830 1985, Evans et al. *J Am Chem Soc* **103** 2127 1981.]

The **4S-(+)-4-isopropyl-3-propionyl-1:3-oxazolidine-2-one derivative** [77877-9-1] **M 185.2, has b 102-106°/0.75mm, d<sub>4</sub><sup>25</sup> 1.094, n<sub>D</sub><sup>20</sup> 1.464,  $[\alpha]_D^{25} +93^\circ$  (c 8.7, CH<sub>2</sub>Cl<sub>2</sub>)**. An auxiliary reagent also used as a chiral ligand in dirhodium (II) complexes [Doyle et al. *J Am Chem Soc* **115** 9968 1993], and in aldol addition reactions [Pridgen et al. *J Org Chem* **58** 5107 1993]. For a reviews on chiral auxiliaries for asymmetric synthesis see Ager et al. *Aldrichimica Acta* **30** 3 1997, and Mukaiyama *Aldrichimica Acta* **29** 59 1996.

**4R-(+)-4-Isopropyl-2-oxazolidinethione [4R-(+)-4-isopropyl-1,3-oxazolidin-2-thione, (4R)-4-(1-methyl-ethyl)-2-oxazolidinethione]** [1217463-35-8] and **4S-(-)-4-isopropyl-2-oxazolidinethione (4S-(-)-4-isopropyl-1,3-oxazolidin-2-thione)** [104499-08-3] **M 145.2, m 48-52°, 51-53°, R-  $[\alpha]_D^{20} +23.2^\circ$ , S-  $[\alpha]_D^{20} -23.2^\circ$ , (c 0.4, CHCl<sub>3</sub>)**. These compounds are Evans type of chiral auxiliaries. The *S*-(-)-enantiomer is synthesised by adding CS<sub>2</sub> (0.9ml, 15mmol) to a solution of *S*-valinol (10mmol [cf 2026-48-4]) in aqueous N Na<sub>2</sub>CO<sub>3</sub> (20ml) and stirring at 100° (bath at 110° under efficient reflux and fume-cupboard) for 15 minutes, cooling to 20° and extracting with CH<sub>2</sub>Cl<sub>2</sub> (2 x 50ml). The extract is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated to dryness and the residue is recrystallised from EtOAc/cyclohexane or EtOAc/hexane. It has UV (EtOH) with  $\lambda_{\max}$  at 244 nm ( $\epsilon$  18,800); the IR has  $\nu_{\max}$  (KBr) at 3160 and 1515 cm<sup>-1</sup>; and the <sup>1</sup>H NMR (300MHz, CDCl<sub>3</sub>) has  $\delta$  at 0.77 (d, 3H, *J* = 6.8Hz), 0.82 (d, 3H, *J* = 6.7Hz), 1.68 (m, 1H), 3.77 (d of t, 1H, *J* = 6.6 and 9.1Hz), 4.23 (d of d, 1H, *J* = 6.6 and 9.1Hz), 4.55 (t, 1H, *J* = 9.1Hz), 9.00 (br s, 1H); and the <sup>13</sup>C NMR (75.5MHz, CDCl<sub>3</sub>) has  $\delta$  at 189.54, 73.53, 62.52, 32.17, 18.04, 17.90. [Delaunay et al. *J Org Chem* **60** 6604 1995, Nagao et al. *J Chem Soc, Perkin Trans I* 2361 1985.] They are selective and efficient chiral auxiliaries [Velazquez & Olivio *Current Org Chem* **6** 303 2000] which can be directly reduced to their corresponding aldehydes and the chiral auxiliary by reductive cleavage with diisobutylaluminium hydride [Crimmins & Chaudhary *Org Lett* **2** 775 2000].

**4R-(+)-4-Isopropyl-2-thiazolidinethione (4R-(+)-4-isopropyl-1,3-thiazolidin-2-thione)** [110199-16-1] and **4S-(-)-4-isopropyl-2-thiazolidinethione (4S-(-)-4-isopropyl-1,3-thiazolidin-2-thione)** [76186-04-4] **M 161.3,**

**m** 66-67°, 67-68°, 69-71°, *R*- $[\alpha]_{\text{D}}^{20} +37^\circ$ , *S*- $[\alpha]_{\text{D}}^{20} -37^\circ$  (**c** 1,  $\text{CDCl}_3$ ). These compounds are efficient *Evans type* of chiral auxiliaries. The *S*-(-)-enantiomer was synthesised by adding  $\text{CS}_2$  (3ml, 50mmol) to a solution of *S*-valinol (10mmol (see [2026-48-4]) in aqueous  $\text{N KOH}$  (50ml) and stirring at 100° (bath at 110° under reflux and efficient fume-cupboard) for 16 hours, cooling to 20° and extracting with  $\text{CH}_2\text{Cl}_2$  (2 x 50ml). The extract is dried ( $\text{Na}_2\text{SO}_4$ ), filtered, evaporated to dryness and the residue is recrystallised from  $\text{CH}_2\text{Cl}_2$  (colourless needles) or  $\text{Et}_2\text{O}$ . *Note* that unlike the preparation of the 1,3-oxazolidine-2-thione above, the preparation of the 1,3-thiazolidine-2-thione required a larger excess of  $\text{CS}_2$ , stronger base and much longer heating time to replace the alcoholic O by S. It has  $^1\text{H NMR}$  (300MHz,  $\text{CDCl}_3$ ) with  $\delta$  at 1.00 (d, 3H,  $J = 7.2\text{Hz}$ ), 1.03 (d, 3H,  $J = 8.5\text{Hz}$ ), 2.01 (m, 1H), 3.32 (d of d, 1H,  $J = 8.2$  and  $11.0\text{Hz}$ ), 3.53 (d of d, 1H,  $J = 8.2$  and  $11.0\text{Hz}$ ), 4.11 (m, 1H), 9.05 (br s, 1H); and the  $^{13}\text{C NMR}$  (75.5MHz,  $\text{CDCl}_3$ ) has  $\delta$  at 200.78, 70.20, 35.73, 31.98, 18.78, 18.18. [Delaunay et al. *J Org Chem* **60** 6604 1995, Nagao et al. *J Chem Soc. Chem Commun* 1418 1985, Nagao et al. *J Org Chem* **51** 2391 1986, McKinnon & Meyer *J Org Chem* **58** 3568 1993.] They are selective and efficient chiral auxiliaries [Velazquez & Olivio *Current Org Chem* **6** 303 2000], and the condensation products can be directly reduced to the corresponding aldehyde and the chiral auxiliary by reductive cleavage with diisobutylaluminium hydride [Crimmins & Chaudhary *Org Lett* **2** 775 2000].

**1*R***(-)-Menthol [natural *l*(-)-, **1*R*,2*S*,5*R***(-)-1-hydroxy-2-isopropyl-5-methylcyclohexane] [2216-51-5] **M** 156.3, **m** 42-45°, 43°, 44-46.5°, 89°/2mm, 100-101°/7mm, 212°/atm,  $d_{25}^{25} 0.89$ ,  $n_{\text{D}}^{25} 1.458$ ,  $n_{\text{D}}^{60} 1.446$ ,  $[\alpha]_{\text{D}}^{20} -50^\circ$  (**c** 10, EtOH),  $[\alpha]_{546}^{18} -58.7^\circ$  (**c** 2, EtOH), and **1*S***(+)-menthol [synthetic *d*(+)-menthol, **1*S*,2*R*,5*S***(+)-1-hydroxy-2-isopropyl-5-methylcyclohexane] [15356-60-2] **m** 43-44°, 103-104°/9mm,  $d_{25}^{25} 0.89$ ,  $n_{\text{D}}^{25} 1.458$ ,  $n_{\text{D}}^{60} 1.446$ ,  $[\alpha]_{\text{D}}^{23} +48^\circ$  (**c** 10, EtOH),  $[\alpha]_{546}^{18} +58.6^\circ$  (**c** 2, EtOH). The natural *l*-isomer is present in peppermint oil and has a strong odour of peppermint, and is sometimes called peppermint camphor. Crystallise menthol from  $\text{CHCl}_3$ , petroleum ether or EtOH/water. It can be sublimed at 40° *in vacuo*, but distillation at 5-10mm is preferable with large quantities. It is best stored under  $\text{N}_2$  in the dark. It is soluble in most organic solvents but is slightly soluble in  $\text{H}_2\text{O}$ . [Barrow & Atkinson *J Chem Soc* 638 1939, *Beilstein* **6** III 133, **6** IV 150.] *l*(-)-Menthol is a very useful resolving agent for acids [Brunel & Buono *J Org Chem* **58** 7313 1993, see also resolution of the Fe complex [12101-02-9] above], and has been used in crystallisation-induced asymmetric transformation of malonate esters [Ihara et al. *J Chem Soc, Chem Commun* 9 1988]. It is a chiral auxiliary that can be recycled [Solladié et al. *Synthesis* 173 1987, Katagiri et al. *J Org Chem* **53** 227 1988]. Similar purification and applications are applicable for *non-natural d*(+)-menthol, with the advantage of producing the optical enantiomers of the products. The *racemic* form **1*R*,2*S*,5*R***(±)-1-hydroxy-2-isopropyl-5-methylcyclohexane (*hexahydro-thymol*) [1490-04-6] **M** 156.3, **m** 28° and 38° (dimorphic), **b** 216.5°/atm,  $d_{30}^{30} 0.8911$ ,  $n_{\text{D}}^{20} 1.4415$ ,  $n_{\text{D}}^{60} 1.4461$ , is obtained by catalytic hydrogenation of thymol [89-83-8] followed by distillation. [Huggett *J Soc Chem Ind* **60** 67 1941, Waters & Beal *J Am Pharm Assoc* **34** 52 1945, *Beilstein* **6** III 137, **6** IV 152.]

## LEWIS AND BRØNSTED/LOWRY ACIDS AND BASES

A few words are warranted here because these terms are frequently used in current literature. The definition of acids (which produce  $\text{H}^+$  ions) and bases (which produce  $\text{HO}^-$  ions) was adequate to explain reactions (e.g. salt formation) in aqueous solutions, and led to the concepts of pH (S.P.L. Sørensen *Biochem Z* **21** 131, 201 1909) and pK (ionisation, cf: Chapter 1 pp 34-36). The definition becomes unsatisfactory when applied to studies of reactions in non-aqueous media, particularly in the catalytic context. Independently, J.N. Brønsted [*Rec Trav Chim Pays Bas* **42** 718 1923, *Chem Rev* **5** 231 1928] and T.M. Lowry [*J Soc Chem Ind* (London) **42** 42 1923] developed the view that an acid is a substance that has a tendency to lose a proton, and a base is one that has a tendency to gain a proton. This led to the understanding of *conjugate* species, e.g.  $\text{R}_3\text{NH}^+$  as potential acids, and defining the equilibrium:  $\text{Base} + \text{H}^+ \rightleftharpoons \text{H-Base}^+$ . They pointed out that the differences in net charge is not as important as the chemical behaviour on which their definition is based. Basicity and acidity do not bear a simple relationship to the net respective charges, and both their properties depend much more on the complex electronic constitutions of the reagents. Later G.N. Lewis (1928 and later work) began with the classical concept of acids and bases, and progressed to a broader definition of an acid as a substance that is able to accept a pair of electrons, and a base as a substance capable of supplying a pair of electrons. This broader definition has been used extensively for reactions in non-aqueous solutions as well as in aqueous solutions. [Note that these concepts originated from extensive studies of catalytic reactions, e.g. mutarotation, hydrolysis, etc]. Thus a

Lewis acid is a substance that is electron deficient (e.g.  $\text{BF}_3$ ), and a Lewis base is a substance that can donate electrons (e.g. amines, phosphines, boranes, ethers, sulfides etc) to form bonds or complexes with Lewis acids.

In the *Bronsted/Lowry definition*, a base donates an electron pair to a proton to form a covalent B—H bond (a positive charge, if involved, will reside on B). In the *Lewis definition*, a base donates a pair of electrons to an electron deficient atom (other than a proton). An acid does not donate a proton but accepts a base to form a new bond. An electron pair is required for forming a covalent or a dative bond.

In the broadest form, an acid is an *electrophile* whereas a base is a *nucleophile*. When these species are regenerated during reactions then they become catalytic. For general reading see Michael B. Smith *Organic Chemistry: An Acid-Base Approach* CRC Press October 2010, ISBN 10: 1420079204, 13: 9781420079203.

## BIBLIOGRAPHY

### Earlier research papers on catalytic metal-mediated carbocyclisations:

**Cobalt-mediated:** Exon & Magnus *J Am Chem Soc* **105** 2477 1983; Billington & Willison *Tetrahedron Lett* **25** 4041 1984; Smit, Gybin, Shashkov, Strychkov, Kyz'mina, Mikaelian, Caple & Swanson *Tetrahedron Lett* **27** 1241 1986; Almansa, Moyano & Serratos *Tetrahedron* **44** 2657 1988; Schore & Rowley *J Am Chem Soc* **110** 5224 1988; Lautens & Crudden *Tetrahedron Lett* **30** 4803 1989; Harvey, Johnson, Ung & Vollhardt *Synlett* 15 1989.

**Gold-Mediated (with Au-NHCs and other ligands):** Ricard & Gagosz *Organometallics* **26** 4704 2007, Echavarren and coworkers *Angew Chem, Int Ed* **44** 6146 2005, Nieto-Oberhuber et al. *J Am Chem Soc* **127** 6168 2005, and **45** 1105 2006, Nieto-Oberhuber *Chem Eur J* **12** 1677 2006.

**Iron-mediated:** Takacs, Anderson, Madhavan, Creswell, Seely & Devroy *Organometallics* **5** 2395 1986; Takacs & Anderson *J Am Chem Soc* **109** 2200 1987; Takacs, Anderson, Creswell & Takacs *Tetrahedron Lett* **28** 5627 1987; Pearson & Zettler *J Am Chem Soc* **111** 3908 1989; Saino, Kogure, Kase & Okamoto *J Organomet Chem* **691** 3129 2006; Takacs, Newsome, Kuehn & Takusagawa *Tetrahedron* **46** 5507 1990.

**Molybdenum-mediated:** Fujimura & Grubbs *J Am Chem Soc* **118** 2499 1996, Alexander, La, Cefalo, Hoveyda & Schrock *J Am Chem Soc* **120** 4041 1998, La, Alexander, Cefalo, Graf, Hoveyda & Schrock *J Am Chem Soc* **120**, 9720 1998, Weatherhead, Houser, Ford, Jamieson, Schrock & Hoveyda *Tetrahedron Lett* **41** 9553 2000,

**Nickel-mediated:** Behr, Freudenberg & Keim *J Mol Cat* **35** 9 1986; Chiusoli *J Organomet Chem* **300** 57 1986; Hernandez & Hoberg *J Organomet Chem* **328** 403 1987; Trost & Tour *J Am Chem Soc* **109** 5268 1987; Tamao, Kobayashi, Ito *J Am Chem Soc* **110** 1286 1988; Trost & Tour *J Am Chem Soc* **110** 5231 1988; Wender, Ihle, Correia *J Am Chem Soc* **110** 5904 1988; Tsuda, Kiyoi, Miyane & Saegusa *J Am Chem Soc* **110** 8570 1988; Wender & Jenkins *J Am Chem Soc* **111** 6432 1989.

**Palladium-mediated:** Stille & Tanaka *J Am Chem Soc* **109** 3785 1987; Trost & Tanoury *J Am Chem Soc* **110** 1636 1988; Abelman & Overman *J Am Chem Soc* **110** 2328 1988; Trost & Lee *J Am Chem Soc* **110** 7255 1988; Trost & Matsuda *J Am Chem Soc* **110** 5233 1988; Trost & Luengo *J Am Chem Soc* **110** 8239 1988; Oppolzer *Angew Chem Int Ed* **28** 38 1989; Trost *Angew Chem Int Ed* **28** 213 1989; Zhang & Negishi *J Am Chem Soc* **111** 3454 1989; Trost & Lee *J Org Chem* **54** 2217 1989; Takacs & Zhu *J Org Chem* **54** 5193 1989; Negishi, Iyer & Rousset *Tetrahedron Lett* **30** 291 1989; Larock, Song, Baker & Gong *Tetrahedron Lett* **29** 2919 1988; Trost *Acc Chem Res* **23** 34 1990; Takacs & Zhu *Tetrahedron Lett* **31** 1117 1990.

**Rhodium-mediated:** James & Young *J Organomet Chem* **285** 321 1985; Funakoshi, Togo, Taura & Sakai *Chem Pharm Bull Jpn* **37** 1776 1989; Taura, Tanaka, Funakoshi & Sakai *Tetrahedron Lett* **30** 6349 1989; Shibata et al. *Tetrahedron* **63** 12853 2007.

**Zirconium-mediated:** Negishi & Takahashi *Synthesis* 1 1988; Buchwald & Nielsen *Chem Rev* **88** 1947 1988; Negishi, Holmes, Tour, Miller, Cederbaum, Swansson & Takahashi *J Am Chem Soc* **111** 3336 1989; Nugent & Taber *J Am Chem Soc* **111** 6435 1989; Yasuda, Okamoto, Matsuoka, Nakamura, Kai, Kanehisa & Kasai *Organometallics* **8** 1139 1989.

### Books

B.C. Gates, H. Knoezinger and F. Jentoft, *Advances in Catalysis*, Academic Press, 2010. ISBN 9780123808523.

- P. Barbaro and F. Liguori, *Heterogenized Homogeneous Catalysts for Fine Chemicals Production*, Springer Verlag, 2010. ISBN 9789048136957.
- M. Beller and C. Bolm (Eds), *Transition Metals for Organic Synthesis*, 2 volumes, 2<sup>nd</sup> edition, Wiley-VCH, 2004. ISBN 3527306137.
- M. Benaglia, *Recoverable and Recyclable Catalysis*, J. Wiley & Sons, 2009. ISBN 9780470681954.
- G.C. Bond, *Metal-Catalysed Reactions of Hydrocarbons*, Springer Verlag, 2010. ISBN 9781441936943.
- G.C. Bond, C. Louis and D.T. Thompson, *Catalysis by Gold*, Imperial College Press, London, 2006. ISBN 9731860946585.
- H. Brunner and W. Zettlmeier, *Handbook of Enantioselective Catalysis with Transition Metal Compounds*, Vol 1, Products and Catalysts and 2, Ligands references, VCH Verlagsgesellschaft, Weinheim, 1993. ISBN 1560818115.
- R.M. Bullock (Ed), *Catalysis without Precious Metals*, Wiley-VCH, 2010. ISBN 9783527323548.
- V. Caprio and J.M.J. Williams, *Catalysis in Asymmetric Synthesis*, J. Wiley & Sons, 2009. ISBN 9781405190916.
- C.S.J. Cazin, *N-Heterocyclic Carbenes in Transition Metal Catalysis and Organocatalysis*, Springer Verlag, 2010. ISBN 9789048128655.
- R. Chauvin and Y. Canac (Eds), *Topics in Organometallic Chemistry V 30, Transition Metal Complexes of Neutral  $\eta^1$ -Carbon Ligands*, Springer, 2010. ISBN 978364207213.
- B. Cornils and W.A. Hermann (Eds), *Applied Homogeneous Catalysis with Organometallic Compounds; a comprehensive handbook in three volumes*, Wiley-VCH Verlag, 2002. ISBN 9783527304349/3527304347.
- B. Cornils, W.A. Hermann, M. Muhler, and C.-H. Wong (Eds), *Catalysis from A to Z. A Concise Encyclopedia*, Wiley-VCH, 2007. ISBN 9783527314386.
- B. Cornils, W.A. Hermann, I. T. Horvath, W. Leitner, S. Mecking, H. Olivier-Bourbiguo, and D. Vogt (Eds), *Multiphase Homogeneous Catalysis, Vol 1 and Vol 2*, Wiley-VCH, 2005. ISBN 9783527307210.
- R. H. Crabtree and D. M. Mingos (Eds-in-Chief), *Comprehensive Organometallic Chemistry, volumes 1—13*, Elsevier, 2007. ISBN 9780080445908.
- M.P. Doyle, M.A. McKervey and T. Ye, *Modern Catalytic Methods for Organic Synthesis with Diazo Compounds: From Cyclopropanes to Ylides*, Wiley, 1998. ISBN 0471135569.
- F.I. Dalko, *Enantioselective Organic Catalysis*, Wiley-VCH, 2007, ISBN 9783527315222.
- A. de Meijere and F. Diedrich (Eds), *Metal-Catalyzed Cross-Coupling Reactions*, 2 volumes, Wiley-VCH, 2004. ISBN 9783527305186/3527305181, 9783527305186/3527305181.
- V. Dragutan, A. Demonceau and I. Dragutan, *Green Metathesis Chemistry*, Springer Verlag, 2009. ISBN 9789048134311, 9789048134328.
- G. Dyker (Ed.), *Handbook of C-H Transformations: Applications in Organic Synthesis*, Vol 1 and 2, Wiley-VCH, 2005. ISBN 9783527310746.

J.J. Eisch and R.B. King, *Organometallic Synthesis Vol 1* library of congress cat no. 65-18428, 1965, *Vol 2* ISBN 0122349504, 1981, *Vol 3* ISBN 0444426078 1986 and *Vol 4* ISBN 044429565, 1988, Academic Press, series ISBN 0444426183.

P.A. Evans (Ed.), *Modern Rhodium-catalysed Organic Reactions*, Wiley-VCH, 2005. ISBN 3527306838.

W.P. Griffith (Ed.), *Ruthenium Oxidation Complexes*, Springer Verlag, 2010. ISBN 9781402093760.

R.H. Grubbs (Ed.), *Handbook of Metathesis*, Vols 1-3, Wiley-VCH, 2003. ISBN 3527306161.

Z. Guan, *Metal Catalysts in Olefine Polymerisation*, Springer Verlag, 2008. ISBN 9783540877509.

L.S. Hegedus and B.C.G. Söderberg, *Transition Metals in the Synthesis of Complex Organic Molecules*, 3<sup>rd</sup> edition, University Science Book, Sausalito, California, 2010. ISBN 9781891389597.

E.N. Jacobsen, A. Pfaltz and H. Yamamoto, *Comprehensive Asymmetric Catalysis*, 3 volumes, Springer Verlag, Berlin. 1999, ISBN 3-540-64336-2; Supplement 1, 2004, ISBN 3-540-00333-9; Supplement 2, 2004, ISBN 3-540-20983-2.

P. Knochel (Ed.), *Handbook of Functionalized Organometallics, Applications in Synthesis*, Wiley-VCH, 2005. ISBN 9783527311316.

P.W.N.M. van Leeuwen, *Homogeneous Catalysis, understanding the art*, Kluwer Academic Publ., London, 2004. ISBN 1402019998.

M. Marafi, S. Anthony and E. Furimsky, *Handbook of Spent Hydroprocessing Catalysts: Regeneration, Rejuvenation, Reclamation, Environment and Safety*, Elsevier, 2010. ISBN 9780444535566.

D.L. Marmaduke, *Progress in Heterogeneous Catalysis*, Nova Science Publ, 2008. ISBN 9781604569810.

N. Mizuno, *Modern Heterogeneous Oxidation Catalysis: Design, Reactions and Characterisation*, Wiley-VCH, 2009. ISBN 9783527318599.

S.-I. Murahashi (Ed), *Ruthenium in Organic Synthesis*, Wiley-VCH, 2004. ISBN 3527306927.

E-I. Negishi (Ed), *Handbook of Organopalladium for Organic Synthesis, Vol 1 and 2*, J, Wiley & Sons, 2002. ISBN 0471315060.

I. Ojima (Ed), *Catalytic Asymmetric Synthesis*, 2<sup>nd</sup> Edn, Vols 1-2, Wiley-VCH, 2000. ISBN 0471298050, 9780471298052.

I. Ojima (Ed.), *Catalytic Asymmetric Synthesis*, 3<sup>rd</sup> Edn, J. Wiley & Sons, 2010. ISBN 9780470175774.

B. Plietker (Ed), *Iron Catalysis*, Springer Verlag, 2011. ISBN 9783642146695.

B. Plietker (Ed), *Iron Catalysis in Organic Chemistry (Reactions and Application)*, Wiley-VCH, 2008. ISBN 9783527319275.

M.L. Prunier, *Catalysis in Organic Reactions*, CRC Press, 2008. ISBN 9781420070767.

G.A. Somorjai and Y. Lin, *Introduction to Surface Chemistry and Catalysis*, J. Wiley & Sons, 2010. ISBN 9780470508237.

P. Roesky, *Molecular Catalysis of Rare-earth Elements*, Springer Verlag, 2010. ISBN 9783642128103.

V. F. Slagt, A. H. M. de Vries, J. G. de Vries and R. M. Kellogg, Practical aspects of carbon cross-coupling reactions using heteroarenes (A Review), *Organic Process Research & Development (OPRD)* **14** 30 2010.

G. Smith and F. Notheisz, *Heterogeneous Catalysis in Organic Chemistry*, Elsevier, 2011. ISBN 9780126516456, 0126516456.

J.J. Spivey and K.M. Dooley (Eds), *Catalysis*, SPR Catalysis, RSC Publ, 2010. ISBN 9781847559517.

C.E. Song, *Cinchona Alkaloids in Synthesis and Catalysis: Ligands, Immobilization and Organocatalysis*, Wiley-VCH, 2009. ISBN 9783527324163.

J.J. Spivey (Series Ed.), *RSC Catalysis Series: E. Furimsky Carbons and Carbon Supported Catalysts in Hydroprocessing*, 2008, ISBN 9780854041435; H. Pellisier, *Chiral Sulfur Ligands*, 2009, ISBN 9781847559241; H. Pellisier, *Recent Developments in Asymmetric Organocatalysis*, 2010, ISBN 9781849730549; J.L. Vicario, D. Badia, L. Carrillo and E. Reyes, *Organocatalytic Enantioselective Conjugate Addition Reactions*, 2010, ISBN 9781849730242; S. Diez-González (ed), *N-Heterocyclic Carbenes*, 2010, ISBN 9781849730426; A. de Klerk and E. Furimsky, *Catalysis in the Refining of Fischer-Tropsch Syncrude*, 2010, ISBN 9781849730808; A. Grabulosa, *P-Stereogenic Ligands in Enantioselective Catalysis*, 2010, ISBN 9781849731232; M. Shi, F. Wang, M-X. Zhao and Y. Wei, *Chemistry of the Morita-Baylis-Hillman Reaction*, 2011, ISBN 9781849731294.

P. Stépnicka (Ed.), *Ferrocenes: Ligands, materials and biomolecules*, J Wiley & Sons, 2008. ISBN 9780470035856.

A. Togni and T. Hayashi, *Ferrocenes: Homogeneous Catalysis, Organic Synthesis and Material Science*, VCH Publ Weinheim 1995. ISBN 3527290486.

P.J. Walsh and M.C. Kozlowski, *Fundamentals of Asymmetric Catalysis*, University Science Book, Sausalito, California, 2008. ISBN 9781891389542.

T.R. Ward, *Bio-inspired Catalysts*, Springer Verlag, 2008. ISBN 9783540877561.

H. Werner, *Landmarks in Organo-transition Metal Chemistry, A Personal View*, Springer Verlag, 2009. ISBN 9780387098470.

Applied Homogeneous Catalysis with Organometallic Compounds: *Transition Metals for Organic Synthesis* 2<sup>nd</sup> Ed, Wiley-VCH Verlag, GmbH (Weinheim) Germany 2002.

# CHAPTER 7

## PURIFICATION OF BIOCHEMICALS AND RELATED PRODUCTS

### INTRODUCTION

Biochemicals are chemical substances produced by living organisms. They range widely in size, from simple molecules such as formic acid and glucose to macromolecules such as proteins and nucleic acids. Their *in vitro* synthesis is often impossibly difficult, and in such cases they are available (if at all) only as commercial tissue extracts which have been subjected to purification procedures of widely varying stringency. The desired chemical may be, initially, only a minor constituent of the source tissue which may vary considerably in its composition and complexity. Recent advances in molecular biology have made it possible to produce substantial amounts of biological materials, which are present in nature in extremely small amounts, by recombinant DNA technology and expression in bacteria, yeast, insect and mammalian cells. The genes for these substances can be engineered such that the gene products, e.g. polypeptides or proteins, can be readily obtained in very high states of purity, and in large amounts if necessary. However, many such products, which are still obtained from the original natural sources, are available commercially and may require further purification.

As a preliminary step the tissue might be separated into phases [e.g. whole egg into white and yolk, blood into plasma (or serum) and red cells], and the desired phase may be homogenised. Subsequent treatment usually comprises filtration, solvent extraction, salt fractionation, ultracentrifugation, chromatographic purification, gel filtration and dialysis. Fractional precipitation with ammonium sulfate gives crude protein species. Purification is finally judged by the formation of a single band of macromolecule (e.g. protein, DNA) on electrophoresis and/or analytical ultracentrifugation. Although these generally provide good evidence of high purity, nonetheless it does not follow that one band under one set of experimental conditions is an absolute indication of homogeneity [D.S. Vodopich and R. Moore, *Biology Laboratory Manual*, McGraw-Hill, 2007, ISBN 9780072995220].

During the past 20 or 30 years a wide range of methods for purifying substances of biological origin have become available. For small molecules (including many sugars and amino acids) reference should be made to Chapters 1 and 2. The more important methods used for large molecules, polypeptides and proteins in particular, comprise:

1. *Centrifugation*. In addition to centrifugation for sedimenting proteins after ammonium sulfate precipitation in dilute aqueous buffer, the technique has been used for fractionation of large molecules in a denser medium or a medium of varying density. By layering sugar solutions of increasing densities in a centrifuge tube, proteins can be separated in a sugar-density gradient by centrifugation. Smaller DNA molecules (e.g. plasmid DNA) can be separated from RNA or nuclear DNA by centrifugation in aqueous cesium chloride (ca 0.975g/ml of buffer) for a long time (e.g. 40 hours at 40,000 x g). The plasmid DNA band appears at about the middle of the centrifuge tube and is revealed by the fluorescent pink band formed by the binding of DNA to ethidium bromide which is added to the CsCl buffer. *Microfuges* are routinely used for centrifugation in Eppendorf tubes (1.2-2ml) and can run up to speeds of 12,000 x g. *Analytical centrifugation*, which is performed under specific conditions in an analytical ultracentrifuge is very useful for determining purity, aggregation of protein subunits and the molecular weight of macromolecules. [D. Rickwood, T.C. Ford and J. Steensgaard (Eds), *Centrifugation: Essential Data Series*, J Wiley & Sons, 1994, ISBN 9780471942719; L.L. Regel and W.R. Wilcox, *Processing by Centrifugation*, Springer, 2001, ISBN 9780306466546; J.M. Graham and D. Rickwood, *Biological Centrifugation*, Springer, 2001, 9781859960370; A. Records and K. Sutherland, *Decanter Centrifugation Handbook*, Elsevier, 2001, ISBN 1856173690].

2. *Gel filtration* with polyacrylamide (mol wt exclusion limit from 3000 to 300,000) and agarose gel (mol wt exclusion limit 0.5 to  $150 \times 10^6$ ) is useful for separating macromolecules. In this technique high-molecular-weight substances are too large to fit into the gel microapertures and pass rapidly through the matrix (with the void volume), whereas low-molecular-weight species enter these apertures and are held there for longer periods of time, being retarded by the column material in the equilibria, relative to the larger molecules. This method is also used for desalting solutions of macromolecules.
- Dry gels* and *crushed beads* are also useful in the gel filtration process. Selective retention of water and inorganic salts by the gels or beads (e.g. Sephadex G-25) results in increased concentration and purity of the protein fraction which moves with the void volume. (See also section on "Gel filtration" in Chapter 1.)
3. *Ion-exchange matrices* are microreticular polymers containing carboxylic acid (e.g. Bio-Rad 70) or phosphoric acid (Pharmacia, Amersham Biosciences, Mono-P) exchange functional groups for weak acidic cation exchangers, sulfonic acid groups (Dowex 50W) for strong acidic cation exchangers, diethylaminoethyl (DEAE) groups for weakly basic anion exchangers and quaternary ammonium (QEAE) groups for strong anion exchangers. The old cellulose matrices for ion exchanges have been replaced by Sephadex, Sepharose or Fractogel which have more even particle sizes with faster and more reproducible flow rates. Some can be obtained in fine, medium or coarse grades depending on particle size. These have been used extensively for the fractionation of peptides, proteins and enzymes. The use of pH buffers controls the strength with which the large molecules are bound to the support in the chromatographic process. Careful standardisation of experimental conditions and similarly the very uniform size distribution of Mono beads have led to high resolution in the purification of protein solutions. MonoQ is a useful strong anion exchanger, and MonoS is a useful strong cation exchanger, whereas MonoP is a weak cation exchanger (Pharmacia, Amersham Biosciences and alternative sources, see Chapter 1). These have been successful with medium pressure column chromatography (HPLC, see below in 7). Chelex 100 binds strongly and removes metal ions from macromolecules. [See sections on "HPLC", "Ion-exchange Resins" and "Ion-exchange Celluloses and Sephadex" in Chapter 1.]
4. *Hydroxylapatite* is used for the later stages of purification of enzymes. It consists essentially of hydrated calcium phosphate which has been precipitated in a specific manner. It combines the characteristics of gel and ionic chromatography. Crystalline hydroxylapatite is a structurally organised, highly polar material which, in aqueous solution (in buffers), strongly adsorbs macromolecules such as proteins and nucleic acids, permitting their separation by virtue of the interaction with charged phosphate groups and calcium ions, as well as by physical adsorption. The procedure therefore is not entirely ion-exchange in nature. Chromatographic separations of singly and doubly stranded DNA are readily achievable, whereas there is negligible adsorption of low-molecular-weight species.
5. *Affinity chromatography* is a chromatographic technique whereby the adsorbant has a particular and specific affinity for one of the components of the mixture to be purified. For example the adsorbant can be prepared by chemically binding an inhibitor of a specific enzyme (which is present in the crude complex mixture) to a matrix (e.g. Sepharose). When the mixture of impure enzyme is passed through the column containing the adsorbant, only the specific enzyme binds to the column. After adequate washing, the pure enzyme can be released from the column by either increasing the salt concentration (e.g. NaCl) in the eluting buffer or adding the inhibitor to the eluting buffer. The salt or inhibitor can then be removed by dialysis, gel filtration (above) or ultrafiltration (see below). [See W.H. Scouten, *Affinity Chromatography: Bioselective Adsorption on Inert Matrices*, J.Wiley & Sons, NY, 1981, ISBN 0471026492; H. Schott, *Affinity Chromatography: Template Chromatography of Nucleic Acids and Proteins*, Marcel Dekker, NY, 1984, ISBN 0824771117; P. Matejtschuk Ed. *Affinity Separations* Oxford University Press 1997 ISBN 0199635501 (paperback); M.A. Vijayalakshmi, *Biochromatography, Theory and Practice*, Taylor & Francis Publ, 2002, ISBN 0415269032; and the section on "Other Types of Chromatography" in Chapter 1.]
6. In the *Isoelectric focusing* of large charged molecules on polyacrylamide or agarose gels, slabs of these are prepared in buffer mixtures (e.g. ampholines) which have various pH ranges along the length of the gel. When a voltage is applied for some time, the buffers arrange themselves on the slabs in respective areas according to their pH ranges (prefocusing). Then the macromolecules are applied near the middle of the slab and allowed to migrate in the electric field until they reach the pH area similar to their isoelectric points and focus at that position. This technique can also be used in a chromatographic mode, *chromatofocusing*, whereby a gel in a column is run (also under HPLC conditions) in the presence of ampholines (narrow or wide pH ranges as required) and the macromolecules are then run through in a buffer. *Capillary electrophoresis* systems in which a current is applied to set the gradient are now available in which the columns are fine capillaries and are used for qualitative and quantitative purposes [See R. Kuhn and S. Hoffstetter-Kuhn, *Capillary Electrophoresis: Principles and Practice*, Springer-Verlag Inc, NY, 1993; P. Camilleri Ed. *Capillary Electrophoresis-Theory and Practice*, CRC Press, Boca Raton, Florida, 1993; D.R. Baker, *Capillary Electrophoresis*, J Wiley & Sons, NY, 1995; P.G. Righetti,

- A. Stoyanov and M. Zhukov, *The Proteome Revisited, Isoelectric Focusing; J. Chromatography Library Vol 63 2001*, Elsevier, ISBN 0444505261. P. Schmitt-Kopplin, *Capillary Electrophoresis: Methods and Protocols*, Humana, 2007, ISBN 9781588295392; J.P. Landers, *Handbook of Capillary & Microchip Electrophoresis and Associated Microtechniques*, CRC Press, Boca Raton, Florida, 2007, ISBN 9780849333293, C. Henry, *Microchip Capillary Electrophoresis*, Humana, 2006, ISBN 9781588292933.] The bands are eluted according to their isoelectric points. Isoelectric focusing standards are available which can be used in a preliminary run in order to calibrate the effluent from the column, or alternatively the pH of the effluent is recorded using a glass electrode designed for the purpose. Several efficient commercial equipment are available for separating proteins on a preparative and semi-preparative scale.
7. *High performance liquid chromatography (HPLC)* is liquid chromatography in which the eluting liquid is sent through the column containing the packing (materials as in 2-6 above, which can withstand higher than atmospheric pressures) under pressure. On a routine basis this has been found useful for purifying proteins (including enzymes) and polypeptides after enzymic digestion of proteins or chemical cleavage (e.g. with CNBr) prior to sequencing (using reverse-phase columns such as  $\mu$ -Bondapak C18). Moderate pressures (50-300psi) have been found most satisfactory for large molecules (FPLC). [See *Scopes Anal Biochem* **114** 8 1981; *High Performance Liquid Chromatography and Its Application to Protein Chemistry*, Hearn in *Advances in Chromatography*, **20** 7 1982; B.A. Bidlingmeyer *Practical HPLC Methodology and Applications*, J Wiley & Sons, NY 1991; L.R. Snyder, J.L. Glajch and J.J. Kirkland *Practical HPLC Method Development*, J Wiley & Sons, NY 1988; ISBN 0471627828; R.W.A. Oliver, *HPLC of Macromolecules: A Practical Approach*, 2nd Edn, Oxford University Press, 1998, T. Hanai, *HPLC: A Practical Guide*, Royal Society of Chemistry (UK), 1999, ISBN 084045155; P. Millner *High Resolution Chromatography*, Oxford University Press, 1999 ISBN 0199636486; see also Chapter 1, Bibliography.]
  8. *Ultrafiltration (UF)* using a filter (e.g. Millipore) can remove water and low-molecular-weight substances without the application of heat. Filters with a variety of molecular-weight exclusion limits not only allow the concentration of a particular macromolecule to be determined, but also the removal (by washing during filtration) of smaller molecular-weight contaminants (e.g. salts, inhibitors or cofactors). This procedure has been useful for changing the buffer in which the macromolecule is present (e.g. from Tris-Cl to ammonium carbonate), and for desalting. Ultrafiltration can be carried out in a stirrer cell (Amicon) in which the buffer containing the macromolecule (particularly protein) is pressed through the filter, with stirring, under argon or nitrogen gas pressure (e.g. 20-60psi). During this filtration process the buffer can be changed. This is rapid (e.g. 2L of solution can be concentrated to a few mls in 1 to 2 hours depending on pressure and filter). A similar application uses a filter in a specially designed tube (Centricon tubes, Amicon) and filtration occurs under centrifugal force in a centrifuge (4-6000rpm at 0°/40min). The macromolecule (usually DNA) then rests on the filter and can be washed on the filter also by centrifugation. The macromolecule is recovered by inverting the filter, placing a conical receiver tube on the same side where the macromolecule rests, filling the other side of the filter tube with eluting solution (usually a very small volume e.g. 100  $\mu$ L), and during further centrifugation this solution passes through the filter and collects the macromolecule from the underside into the conical receiver tube. With the development of polymeric and ceramic nanofilters use can be made of nanofiltration (NF) in which particles or molecules of less than 2nm can be held back. This is to be compared with UF where the size limit is between 2nm and 0.1 $\mu$ m (see Chapter 8).
  9. *Partial precipitation* of a protein in solution can often be achieved by controlled addition of a strong salt solution, e.g. ammonium sulfate. This is commonly the first step in the purification process. Its simplicity is offset by possible denaturation of the desired protein and the (sometimes gross) contamination with other proteins. It should therefore be carried out by careful addition of small aliquots of the powdered salt or concentrated solution (below 4°, with gentle stirring) and allowing the salt to be evenly distributed in the solution before adding another small aliquot. Under carefully controlled conditions and using almost pure protein, it is sometimes possible to obtain the protein in crystalline form suitable for X-ray analysis (see below).
  10. *Dialysis*. This is a process by which small molecules, e.g. ammonium sulfate, sodium chloride, are removed from a solution containing the protein or DNA using a membrane which is porous to small molecules. The solution (e.g. 10ml) is placed in a dialysis bag or tube tied at both ends, and stirred in a large excess of dialysing solution (e.g. 1.5 to 2 L), usually a weak buffer at ca 4°. The dialysing buffer is replaced with fresh buffer several times, e.g. four times in 24 hours. This procedure is similar to ultrafiltration (above) and allows the replacement of buffer in which the protein, or DNA, is dissolved. It is also possible to concentrate the solutions by placing the dialysis tube or bag in Sephadex G25 which allows the passage of water and salts from the inside of the bag thus concentrating the protein (or DNA) solution. Dialysis tubing is available from various distributors, but "Spectra/por" tubing (from Spectrum Medical Industries, Inc, LA) is particularly effective because it retains macromolecules and allows small molecules to dialyse out very rapidly, thus reducing dialysing time considerably. This procedure is used when the buffer has to be changed so as to be compatible with the next purification or storage step, e.g. when the protein (or DNA) needs to be stored frozen in a particular buffer over

extended periods. UF and NF can also serve this purpose whereby the solvent can be completely replaced by washing with an alternative solvent.

11. *Gel Electrophoresis*. This is becoming a more commonly used procedure for purifying proteins, nucleic acids, nucleoproteins, polysaccharides and carbohydrates. The gels can be "electroblotted" onto membranes, and the modern procedures of identifying, sequencing (proteins and nucleic acids) and amplifying (nucleic acids) on sub-micro scales have made this technique a very important one. See below for polyacrylamide gel electrophoresis (PAGE). [D. Patel *Gel Electrophoresis*, J.Wiley-Liss, Inc., 1994; P. Jones and D. Rickwood, *Gel Electrophoresis: Nucleic Acids*, J. Wiley and Sons, 1999 (paperback) ISBN 0471960438; D.M. Gersten and D. Gersten, *Gel Electrophoresis: Proteins*, J. Wiley and Sons, 1996, ISBN 0471962651; R. Westermeier *Electrophoresis in Practice*, 4th Edn, Wiley-VCH Publishing, 2004 ISBN 9783527311811].
12. *Crystallisation*. The ultimate in purification of proteins or nucleic acids is crystallisation. This involves very specialised procedures and techniques and is best left to the experts in the field of X-ray crystallography who can provide a complete picture of the structure of these large molecules. [A. Ducruix and R. Giegé Eds, *Crystallisation of Nucleic Acids and Proteins: A Practical Approach*, 2nd Edition, 2000, Oxford University Press, ISBN 0199636788 (paperback); T.L. Blundell and L.N. Johnson *Protein Crystallisation*, Academic Press, NY, 1976; A. McPherson *Preparation and Analysis of Protein Crystals*, J.Wiley & Sons, NY, 1982; A. McPherson, *Crystallisation of Biological Macromolecules*, Cold Spring Harbour Laboratory Press, 2001 ISBN 0879696176, see also Bibliography in Chapter 1.]

Other details of the above can be found in Chapters 1 and 2 which also contain relevant references.

Several illustrations of the usefulness of the above methods are given in the *Methods Enzymol* series (Academic Press) in which 1000-fold purifications or more have been readily achieved. In applying these sensitive methods to macromolecules, reagent purity is essential. It is disconcerting, therefore, to find that some commercial samples of the widely used affinity chromatography ligand Cibacron Blue F3GA contained this dye only as a minor constituent. The major component appeared to be the dichlorotriazinyl precursor of this dye. Commercial samples of Procion Blue and Procion Blue MX-R were also highly heterogeneous [Hanggi and Cadd *Anal Biochem* **149** 91 1985]. Variations in composition of sample dyes can well account for differences in results reported by different workers. The purity of substances of biological origin should therefore be checked by one or more of the methods given above. Water of high purity should be used in all operations. Double glass distilled water or water purified by a MilliQ filtration system (see Chapter 2) is most satisfactory.

**Brief general procedures for the purification of polypeptides and proteins.** Polypeptides of up to *ca* 1-2000 (10-20 amino acid residues) are best purified by reverse phase HPLC. The desired fractions that are collected are either precipitated from solution with EtOH or lyophilised. The purity can be checked by HPLC and identified by microsequencing (1-30 picomoles) to ascertain that the correct polypeptide was in hand. Polypeptides larger than these are sometimes classified as proteins and are purified by one or more of the procedures described above. The purification of enzymes and functional proteins which can be identified by specific interactions is generally easier to follow because enzyme activities or specific protein interactions can be checked (by assaying) after each purification step. The commonly used procedures for purifying soluble proteins involve the isolation of an aqueous extract from homogenised tissues or extracts from ruptured cells from microorganisms or specifically cultured cells, for example, by sonication, freeze shocking or passage through a small orifice under pressure. Contaminating nucleic acids are removed by precipitation with a basic protein, e.g. protamine sulfate. The soluble supernatant is then subjected to fractionation with increasing concentrations of ammonium sulfate. The required fractions are then further purified by the procedures described in sections 2-9 above. If an affinity adsorbant has been identified, then affinity chromatography can provide an almost pure protein in one step sometimes even from the crude extract. The rule of thumb is that a solution with a protein concentration of 1mg/ml has an absorbance  $A_{1cm}$  at 280nm of 1.0 units. Membrane-bound proteins are usually insoluble in water or dilute aqueous buffer and are obtained from the insoluble fractions, e.g. the microsomal fractions from the  $>100,000 \times g$  ultracentrifugation supernatant. These are solubilised in appropriate detergents, e.g. Mega-10 (nonionic), Triton X-100 (ionic) detergents, and purified by methods 2 to 8 (previous section) in the presence of detergent in the buffer used. They are assayed also in the presence of detergent or membrane lipids.

The purity of proteins is best checked by *polyacrylamide gel electrophoresis* (PAGE). The gels are either made or purchased as pre-cast gels and can be with uniform or gradient gel composition. Proteins are applied onto the gels *via* wells set into the gels or by means of a comb, and travel along the gel surface by means of the current applied to the gel. When the buffer used contains sodium dodecylsulfate (SDS), the proteins are denatured and the denatured proteins (e.g. as protein subunits) separate on the gels mainly according to their molecular sizes. These can be identified by running marker proteins, with a range of molecular weights, simultaneously on a track alongside the proteins under study. The protein bands are visualised by fixing the gel (20% acetic acid) and staining with Coomassie blue followed by silver staining if higher sensitivity is required. An Amersham-Pharmacia “Phast Gel Electrophoresis” apparatus, or related equipment, is very useful for rapid analysis of proteins. It uses small pre-cast polyacrylamide gels (two gels can be run simultaneously) with various uniform or gradient polyacrylamide concentrations as well as gels for isoelectric focusing. The gels are usually run for 0.5-1.5 hours and can be stained and developed (1-1.5 hours) in the same apparatus. The equipment can be used to “electroblot” the protein bands onto a membrane from which the proteins can be isolated and sequenced or subjected to antibody or other identification procedures. It should be noted that all purification procedures are almost always carried out at *ca* 4° in order to avoid denaturation or inactivation of the protein being investigated.

There has been considerable necessity for, and interest in, the study of **Proteomics**. This involves the identification, quantitation and isolation of all the proteins produced by a cell or organism at a particular point in time. It provides information on the expression of **all** the proteins produced by particular cells at a desired stage of the cell's development, maturity, activation or condition. A sophisticated apparatus for this purpose is a flat bed polyacrylamide gel which is run electrophoretically in one direction according to the extent of polymerization of the acrylamide and then run at right angles along a pH gradient (isoelectric focusing). Hundreds of polypeptides and proteins are thus separated, collected and identified by various other techniques such as LC-MS-MS, capillary electrophoresis etc (T. Palzkill, *Proteomics*, Springer, 2001, ISBN 0792375653; T.D. Veenstra and R.D. Smith *Proteome Characterization and Proteomics*, Academic Press, 2003, ISBN 9780120342655; R. Westermeier, T. Naven and H-R. Höpker, *Proteomics in Practice: A Guide to Successful Experimental Design*. J. Wiley & Sons, 2008, ISBN 9783527319411; the *Journal of Proteomics* (ISSN: 1874-3919), which is the official journal of *The European Proteomics Association* published by Elsevier, has been running for several years and is but one of the many journals on Proteomics and Bioinformatics that are available and can be viewed on the internet; see also Bibliography in Chapter 1).

Another rapidly developing field is **metabolomics** where metabolites are screened and identified in the normal and diseased cell at specific time intervals. These cannot be identified from studies of *genomics*, *transcriptomics* or *proteomics*. The subject is now possible because of the highly improved power of HPLC, GC, MS, NMR and the interfacing of these instruments. Thus *metabolones* can be mapped for various biological systems (plant and animal). Publications that are available include the *Metabolomics Journal* published by SpringerLink and started in 2005, and *Journal of Metabolomics and Systems Biology (JMSB)*, published by Academic Journals which started in 2011. These periodicals publish original papers, reviews and conference reports.

Anyone contemplating the purification of a protein is referred to: Professor R.K. Scopes's monograph *Protein Purification*, 3rd Edn, Springer-Verlag, New York, 1994, ISBN 0387940723; M.L. Ladisch Ed. *Protein Purification - from Molecular Mechanisms to Large-scale Processes*, American Chemical Society, Washington DC, 1990; E.L.V. Harris and S. Angal, *Protein Purification Applications - A Practical Approach*, IRL Press, Oxford, 1990; J.C. Janson and L. Rydén, *Protein Purification - Principles, High Resolution Methods and Applications*, VCH Publ. Inc., 1989; ISBN 0895731223, Satinder Ahja *Handbook of Bioseparations*, Academic Press, 2000, ISBN 0120455404; S.M. Wheelwright, *Protein Purification: Design and Scale up of Downstream Processing*, J Wiley & Sons, 1994, references in the bibliography in Chapter 1, and selected volumes of *Methods Enzymol*, e.g. M.P. Deutscher (Ed), *Guide to Protein Purification*, *Methods Enzymol*, Academic Press, Vol **182** 1990, ISBN 0121820831; M.A. Vijayalakshmi, *Biochromatography, Theory and Practice*, Taylor & Francis Publ, 2002, ISBN 0415269032; J.S. Davies, *Amino Acids, Peptides and Proteins Vol 32* 2001, A Specialist Periodical Report, Royal Society of Chemistry, ISBN 0854042326; S. Roe, *Protein Purification Techniques: A Practical Approach*, 2<sup>nd</sup> Edn, Oxford University Press, 2001, ISBN 0199636737; J.M. Walker (Ed) *The Protein Protocols Handbook*, 3<sup>rd</sup> Edn (Springer Protocols Handbooks) Humana Press, 2009, ISBN 978-1-60327-474-6. e-ISBN 978-1-59745-198-7; T. Palmer, *Enzymes, Biochemistry, Biotechnology, Clinical Chemistry*, Horwood Publishing, 2001, ISBN 1898563780. For a comprehensive treatise of many volumes see *Springer Handbook of*

*Enzymes* D. Schonburg & I. Schonburg Eds (A. Chang co-Ed) Springer-Verlag, Berlin, Heidelberg, 2003-onwards (with 39 volumes in 2011) <<http://www.springer.de>>

**Brief general procedures for purifying DNA.** Oligo-deoxyribonucleotides (up to *ca* 60-mers) are conveniently purified by HPLC (e.g. using a Bio-Rad MA7Q anion exchange column and a Rainin Instrument Co, Madison, Dynamax-300A C<sub>8</sub> matrix column) and used for a variety of molecular biology experiments. Plasmid and chromosomal DNA can be isolated by centrifugation in cesium chloride buffer (see section 1, Centrifugation above), and then re-precipitated with 70% ethanol at -70° (18 hours), collected by centrifugation (microfuge) and dried in air before dissolving in TE (10mM TrisHCl, 1mM EDTA pH 8.0). The DNA is identified on an Agarose gel slab (0.5 to 1.0% DNA grade in 45mM Tris-borate + 1mM EDTA or 40mM Tris-acetate + 1mM EDTA pH 8.0 buffers) containing ethidium bromide which binds to the DNA and under UV light causes it to be visualised as pink fluorescent bands. Marker DNA (from  $\lambda$  phage DNA cut with the restriction enzymes Hind III and/or EcoRI ) with bands running from 72 to 353 base-pairs (bp) are run in a parallel track in order to estimate the size of the unknown DNA. Various other DNA markers are commercially available such as the step ladder ranging from 50bp to 800bp with bands at 50bp intervals, and the step ladder with bands ranging from 100bp to 4000bp with bands at 200bp intervals. The DNA can be isolated from the band on the gel by transfer onto nitro-acetate paper (e.g. NA 45) electrophoretically, by binding to silica or an ion-exchange resin, extracted from these adsorbents and precipitated with ethanol. The DNA pellet is then dissolved in TE buffer and its concentration determined. A solution of duplex DNA (or RNA) of 50 $\mu$ g/ml gives an absorbance of 1.0 unit at 260nm/1cm cuvette (single-stranded DNA or RNA gives a value of 1.3 absorbance units). DNA obtained in this way is suitable for molecular cloning.

Recombinant and chemically synthesised DNA and RNA are now routinely separated and purified by HPLC, and their structures are confirmed by sequencing an aliquot. A variety of commercially available HPLC systems are now available, and the desired system can be selected from them.

Brief mention must be made of the tremendous advances that have been made in recent years in the fields of DNA, RNA gene sequencing and synthesis. The development of instrumentation and analysers by the Illumina Company ([www.illumina.com/](http://www.illumina.com/)) and the “*Ion torrent*” semiconductor sequencing of DNA and RNA, using an ion PGM (personal genome machine) sequencer (see; <http://lifetech-it.hosted.jivesoftware.com/index.jspa>) have made it possible to sequence complete genomes in a matter of weeks or less.

For experimental details on the isolation, purification and manipulation of DNA and RNA the reader is referred to: J. Sambrook, E.F. Fritsch and T. Maniatis, *Molecular Cloning-A Laboratory Manual*, 2nd Edn (3 volumes), Cold Spring Harbor Laboratory Press (CSHL Press) NY, 1989, ISBN 0879693096 (paperback); P.D. Darbre, *Basic Molecular Biology: Essential Techniques*, J. Wiley and Sons, 1998, ISBN 0471977055; J. Sambrook and D.W. Russell, *Molecular Cloning-A Laboratory Manual*, 3rd Edn (3 volumes), Cold Spring Harbor Laboratory Press, NY, 2001, ISBN 0079695773, ISBN 9780879695774 (paperback), ISBN 0079695765 (cloth bound); J. Sambrook and D.W. Russell, *The Condensed Protocols for Molecular Cloning: A Laboratory Manual*, CSHL Press, 2006, ISBN 9780879697716, also available on line; M.A. Vijayalakshmi, *Biochromatography, Theory and Practice*, Taylor & Francis Publ, 2002, ISBN 0415269032; A. Travers and M. Buckle, *DNA-Protein Interactions: A Practical Approach*, Oxford University Press, 2000, ISBN 0199636915 (paperback); R. Rapley and D.L. Manning Eds *RNA: Isolation and Characterisation Protocols*, Humana Press 1998 ISBN 0896034941; R. Rapley, *The Nucleic Acid Protocols Handbook*, Humana Press 2000 ISBN 0896038416 (paperback).

This chapter lists some representative examples of biochemicals and their origins, a brief indication of key techniques used in their purification, and literature references where further details may be found. Simpler low-molecular-weight organic compounds, particularly those that may have been prepared by chemical syntheses, e.g. acetic acid, will be found in Chapter 4. Only a small number of enzymes and proteins are included because of space limitations. The purification of the ones that have been included has been described only briefly. The reader is referred to comprehensive texts such as the *Methods Enzymology* (Academic Press, see below) series which currently runs to volume 489 in 2011, and *The Enzymes* (3rd Edn, Academic Press) which runs to 22 volumes for methods of preparation and purification of proteins and enzymes. Leading references on proteins will be found in *Advances in Protein Chemistry* which was incorporated with *Advances in Structural Biology* (84 volumes (2011), Elsevier Inc) and on enzymes will be found in *Advances in Enzymology* which then became *Advances in Enzymology and Related Areas of Molecular Biology*, J Wiley & Sons (up to volume 77 in 2011). The *Annual Reviews of Biochemistry* (Annual Reviews Inc. Patlo Alto California) is also an excellent source of

key references to the up-to-date information on known and new natural compounds, from small molecules, e.g. enzyme cofactors to proteins and nucleic acids. See also the *Springer Handbook of Enzymes* cited above.

**Abbreviations** of titles of periodical are defined as in the Chemical Abstracts Service Source Index (CASSI).

**Ionisation constants** of ionisable compounds are given as **pK** values (published from the literature) and refer to the **pKa** values at room temperature (~15°C to 25°C). The values at other temperatures are given as superscripts, e.g. **pK<sup>25</sup>** for 25°C. Estimated values are entered as **pK<sub>Est(1)</sub>** ~ (see section on “Ionisation Constants” in Chapter 1 for further information).

**Benzene**, which has been used as a solvent successfully and extensively in the past for reactions and purification by chromatography and crystallisation is now considered a **very dangerous substance**, so it has to be used with extreme care. We emphasise that an alternative solvent to benzene (e.g. toluene, toluene-petroleum ether, or a petroleum ether to name a few) should be used first. However, if benzene has to be used then all operations have to be performed in well-ventilated fumehoods and precautions taken to avoid inhalation and contact with skin and eyes. Whenever benzene is mentioned in the text, an asterisk e.g. \*C<sub>6</sub>H<sub>6</sub> or \*benzene, is inserted to remind the user that special precaution should be adopted.

*Selected Amino acids and peptides, Proteins Enzymes DNA and RNA, Carotenoids, Carbohydrates, Steroids, and Miscellaneous Compounds (which include Biochemical reagents, Cofactors, Coenzymes and Vitamins) are collected in the separate respective sections of this chapter.*

## AMINO ACIDS and PEPTIDES

This section includes amino acid derivatives and related compounds.

***N*-Acetyl-L-alaninamide** [15962-47-7] **M 130.2, m 162°**. Crystallise the amide repeatedly from EtOH/diethyl ether. The (±)-isomer crystallises from H<sub>2</sub>O and has **m 157-158°**. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 1838 1961, de Jong *Rec Trav Chim Pays Bas* **19** 288 1900, Fischer & Otto *Chem Ber* **36** 2106 1903, *Beilstein* **4** H 295.]

***N*-Acetyl-β-alanine** [3025-95-4] **M 127.2, m 78.3-80.3°, pK<sup>25</sup> 4.45**. The β-alanine crystallises from acetone. [King & King *J Am Chem Soc* **78** 1089 1956, *Beilstein* **4** IV 2526, 2548.]

***N*-Acetyl-L-alanyl-L-alaninamide** [30802-37-0] **M 201.2, m 250-251°**. Crystallise the dipeptide derivative repeatedly from EtOH/diethyl ether.

***N*-Acetyl-L-alanyl-L-alanyl-L-alaninamide** [29428-34-0] **M 272.3, m 295-300°**. Crystallise the tripeptide derivative from MeOH/diethyl ether.

***N*-Acetyl-L-alanylglycinamide** [76571-64-7] **M 187.2, m 148-149°**. Crystallise the dipeptide derivative repeatedly from EtOH/diethyl ether.

**Acetyl-α-amino-*n*-butyric acid** [34271-24-4] **M 145.2, pK<sup>25</sup> 3.72**. Crystallise the acid twice from water (charcoal) and dry it in air [King & King *J Am Chem Soc* **78** 1089 1956].

**Acetylcarnitine chloride (2-acetoxy-3-carboxy-*N,N,N*-trimethylpropanamine HCl)** [*S*(*D*)- 5080-50-2, *R*(*L*)- 5061-35-8, *RS* 2504-11-2] **M 239.7, m 181°, 187°(corr, dec), 197°(dec), [α]<sub>D</sub><sup>25</sup> -28° (c 2, H<sub>2</sub>O) for *S*-isomer, pK<sup>25</sup> 3.6**. Recrystallise the chloride from isopropanol. Dry it over P<sub>2</sub>O<sub>5</sub> under high vacuum. The *S*-betaine crystallises from EtOH/Et<sub>2</sub>O with **m 145°(dec)** and is hygroscopic; it has [α]<sub>D</sub><sup>20</sup> -19.5° (c 6, H<sub>2</sub>O). [Krimberg & Wittandt *Biochem Z* **251** 231 1932, Strack et al. *Z Physiol Chem* **238** 191 1936, *Beilstein* **4** III 1630, 1632.]

***R*(-)-*N*-Acetyl-L-cysteine methyl ester** [7652-46-2] **M 177.2, m 71-78°, 80°, [α]<sub>D</sub><sup>20</sup> -24.0° (c 1, MeOH)**. The ester is purified by converting into the cuprous mercaptide which is decomposed by dilute H<sub>2</sub>SO<sub>4</sub>, extracted into Et<sub>2</sub>O, dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated and the residue recrystallised from H<sub>2</sub>O containing a little AcOH. The crystals are dried in a vacuum. These operations should be carried out in an inert atmosphere (N<sub>2</sub> or argon) to avoid oxidation to the disulfide cystin ester. Note that the cuprous salt is only stable when it is dry, but is readily oxidised when wet. It has been used as a sulfur transfer agent [Gilman & Spero *Tetrahedron Lett* **33** 1751 1993]. [Pirie *Biochem J* **25** 618 1931, *Beilstein* **4** III 1607.]

***N*-Acetylglutamic acid** [1188-37-0] **M 189.2, m 185° (*RS*), 201° (*S*), [α]<sub>D</sub><sup>25</sup> -16.6° (in H<sub>2</sub>O), [α]<sub>D</sub><sup>25</sup> -5.6° (c 4, MeOH) for *S*-enantiomer, pK<sub>Est(1)</sub> ~3.4, pK<sub>Est(2)</sub> ~4.3**. A likely impurity is glutamic acid. Crystallise it from boiling water. It inhibits *N*-acetyl-L-glutamate synthase. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 1948 1961, Shigesada & Tatibana *Eur J Biochem* **84** 285 1978, Coude *Biochem Biophys Res Commun* **102** 1016 1981, *Beilstein* **4** IV 3047.]

***N*-Acetylglycinamide** [2620-63-5] **M 116.1, m 139-139.5°**. Crystallise the amide repeatedly from EtOH/Et<sub>2</sub>O. Dry it in a vacuum desiccator over KOH. [Davis & Levy *J Chem Soc* 3485 1951, Fischer & Otto *Chem Ber* **36** 2106 1903, *Beilstein* **4** IV 2401.]

***N*-Acetylglycine** [543-24-8] **M 117.1, m 206-208°, pK<sub>1</sub><sup>25</sup> -1.92, pK<sub>2</sub><sup>25</sup> 3.69**. *N*-Acetylglycine is treated with acid-washed charcoal and recrystallised three times from water or EtOH/Et<sub>2</sub>O and is dried *in vacuo* over KOH [King & King *J Am Chem Soc* **78** 1089 1956]. [*Beilstein* **4** IV 2399.]

***N*-Acetylglycyl-L-alaninamide** [34017-20-4] **M 175.2**. Crystallise the dipeptide derivative repeatedly from EtOH/Et<sub>2</sub>O. Dry it in a vacuum desiccator over KOH.

***N*-Acetylglycylglycinamide** [27440-00-2] **M 173.2, m 207-208°**. Crystallise the dipeptide derivative repeatedly from EtOH/Et<sub>2</sub>O. Dry it in a vacuum desiccator over KOH.

***N*-Acetylglycylglycylglycinamide** [35455-24-4] **M 230.2, m 253-255°**. Crystallise the tripeptide derivative repeatedly from EtOH/Et<sub>2</sub>O. Dry it in a vacuum desiccator over KOH.

***N*-Acetylhistidine (H<sub>2</sub>O)** [39145-52-3] **M 171.2, m 148° (RS), 169° (S), [α]<sub>D</sub><sup>25</sup> +46.8° (c 1, H<sub>2</sub>O) for *S*-enantiomer**. A likely impurity is histidine. Crystallise it from water, then 4:1 acetone/water. [Marshall et al. *J Am Chem Soc* **78** 4636 1956, Bergmann & Zervas *Biochem Z* **203** 280 1928, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1990 1961, Beilstein **25** IV 4359.]

***N*-Acetyl-*RS*-homocysteine thiolactone (Citolone)** [1195-16-0] [17896-21-8 for ±] **M 159.2, m 110°, 109-111°, 111.5-112.5°**. Dry Citolone in a vacuum desiccator. It recrystallises from toluene as needles. It is a ninhydrin -ve substance which gives a "slow" nitroprusside test. It has λ<sub>max</sub> at 238nm (ε 4,400 M<sup>-1</sup>cm<sup>-1</sup>); and ν<sub>max</sub> (nujol) 1789s and 851ms cm<sup>-1</sup>. [Benesch & Benesch *J Am Chem Soc* **78** 1597 1956, cf Laliberté *J Chem Soc* 2756 1963.]

***N*-Acetyl-L-leucinamide (2-acetamido-4-methylvaleramide)** [28529-34-2] **M 177.2, m 202°**. Recrystallise it from aqueous EtOH or CHCl<sub>3</sub>/petroleum ether (b 40-60°). [Gränacher *Helv Chim Acta* **8** 216 1925, Beilstein **4** II 864, for L, and Bergmann et al. *Justus Liebigs Ann Chem* **449** 301 1926, Beilstein **4** II 877 for DL.]

***N*-Acetyl-L-methionine** [65-82-7] **M 191.3, m 103.5-104.5°, 104°, [α]<sub>D</sub><sup>25</sup> -24.5° (c 1, in H<sub>2</sub>O), pK<sub>Est</sub> ~3.4**. Crystallise *N*-acetyl-L-methionine from Me<sub>2</sub>CO, H<sub>2</sub>O or EtOAc. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. Its solubility at 25° in H<sub>2</sub>O is 30.7%, and in Me<sub>2</sub>CO it is 29.5%. [Mitzi & Schueter *Biochim Biophys Acta* **27** 168 1958, Birnbaum et al. *J Biol Chem* **194** 455 1952, Beilstein **4** IV 3206.]

**Acetylmethionine nitrile** [538-14-7] **M 172.3, m 44-46°**. Crystallise the nitrile from diethyl ether. [Catch et al. *J Chem Soc* 1611 1947, Beilstein **4** III 1654.]

***N*-Acetyl-*N'*-methyl-L-alaninamide** [19701-83-8] **M 144.2, m 181.2-182° (sealed tube), [α]<sub>D</sub><sup>25</sup> -51.1° (c 2, EtOH)**. Crystallise the amide from EtOAc/Et<sub>2</sub>O, then from EtOH and Et<sub>2</sub>O. Also recrystallise it twice by dissolving ~2.5g in hot 200ml of toluene and cooling. It sublimes at ~170°, so its **m** is measured in a sealed tube. [Applewhite & Niemann *J Am Chem Soc* **81** 2212 1959, Beilstein **4** IV 2500.]

***N*-Acetyl-*N'*-methylglycinamide** [7606-79-3] **M 130.2, m 157.5-158°**. Recrystallise the amide from EtOH/Et<sub>2</sub>O mixture. Also recrystallise it twice from EtOAc/EtOH (16:1) and once from EtOAc. [Applewhite & Niemann *J Am Chem Soc* **81** 2212 1959.]

***N*-Acetyl-*N'*-methyl-L-leucine amide** [32483-15-1] **M 186.3, m 165.3-166.8° (sealed tube), [α]<sub>D</sub><sup>25</sup> -33.9° (c 1, H<sub>2</sub>O)**. Recrystallise the amide from EtOAc, EtOH/hexane mixture or toluene/hexane mixture. [Applewhite & Niemann *J Am Chem Soc* **81** 2212 1959.]

***N*-Acetyl-L-phenylalanine** [2018-61-3] **M 207.2, m 170-171°, 174-175°, [α]<sub>D</sub><sup>25</sup> +47.5° (c 4, EtOH), +52.5° (c 2, EtOH), (DL) m 152.5-153°, pK<sub>Est</sub> ~3.5**. *N*-Acetyl-L-phenylalanine is recrystallised from H<sub>2</sub>O, 20% MeOH/H<sub>2</sub>O, or CHCl<sub>3</sub>; dry and store it at 4°. The (*DL*)-isomer crystallises from H<sub>2</sub>O, Me<sub>2</sub>CO, EtOAc, or CHCl<sub>3</sub> with **m** 152-154° and the solubilities in w% at 25° are 0.73 (H<sub>2</sub>O), 4.3 (Me<sub>2</sub>CO), 0.79 (EtOAc) and 0.34 (CHCl<sub>3</sub>) [Kerr & Niemann *J Org Chem* **23** 893 1958, Overby & Ingersoll *J Am Chem Soc* **73** 3363 1951, L: Fu et al. *J Am Chem Soc* **76** 6057 1954, Bender & Glasson *J Am Chem Soc* **81** 1591 1959]. [Beilstein **14** I 238, **14** IV 1575.]

***N*-Acetyl-L-phenylalanine ethyl ester** [2361-96-8] **M 235.3, m 93-94°**. Crystallise the ester from aqueous

EtOH or H<sub>2</sub>O. [Izumiya & Fruton *J Biol Chem* **218** 59 1956.]

**N-Acetyltryptophan** [87-32-1] **M 246.3**, **m 206°**, **207-208° (RS)**, **pK<sub>Est</sub> ~3.8**, [1218-34-4] **m 188°, 189.5-190.5° (S)**, **[α]<sub>D</sub><sup>25</sup> +30.1° (aqueous NaOH)**, **+71.5° (dioxane/aqueous HCl)**. A likely impurity is tryptophan. Crystallise it from EtOH by adding water. [Cowgill *Biochim Biophys Acta* **200** 18 1970, DL: Berg *J Biol Chem* **100** 79 1933, *Beilstein* **22/14** V 40-50.]

**N-Acetyl-L-valine amide** [37933-88-3] **M 158.2**, **m 275°**. Recrystallise the amide from CH<sub>3</sub>OH/Et<sub>2</sub>O.

**Alamethicin (from *Tricoderma viridae*)**. [27061-78-5] **M 1964.3**, **m 259-260°, 275-270°**, **[α]<sub>D</sub><sup>22</sup> -45° (c 1.2, EtOH)**, **pK<sub>25</sub> 6.04 (aqueous EtOH)**. Recrystallise alamethicin from MeOH. [Panday et al. *J Am Chem Soc* **99** 8469 1977.] The *acetate* [64918-47-4] has **m 195-180°** from MeOH/Et<sub>2</sub>O, and the *acetate-methyl ester* [64936-53-4] has **m 145-140°** from aqueous MeOH.

**α-Alanine (RS)** [302-72-7] **M 89.1**, **m 295-296°**, **(S) [56-41-7] m 297°(dec)**, **[α]<sub>D</sub><sup>15</sup> +14.7° (in 1M HCl)**, **(R) [338-69-2] m 289-291°(dec)**, **[α]<sub>D</sub><sup>15</sup> -14.1° (c 0.9, 1M HCl)**, **pK<sub>1</sub><sup>25</sup> 2.34**, **pK<sub>2</sub><sup>25</sup> 9.87**. Crystallise alanine from H<sub>2</sub>O or aqueous EtOH, i.e. crystallise it from 25% EtOH in water, or recrystallise it from 62.5% EtOH, wash it with EtOH and dry it to constant weight *in vacuo* over P<sub>2</sub>O<sub>5</sub>. *RS-α-alanineamide* [20108-77-4] has **m 62°** (from CHCl<sub>3</sub>), **pKa 8.02**, the *hydrochloride* [80222-96-4] has **m 173°** (from EtOH) and the *acetate salt* has **m 136-137°**. *S-α-alanine methyl ester hydrochloride* [2491-20-5] has **m 109-111°**, **[α]<sub>D</sub><sup>15</sup> +8.0° (c 1.6, MeOH)**, and *S-α-alanine N-methylamide* [7324-05-2] has **m 78°** (hygroscopic, also **m 72°** reported, from CHCl<sub>3</sub>). [Gutter & Kegeles *J Am Chem Soc* **75** 3893 1953, Walsh *J Biol Chem* **264** 2394 1989.] 2,2'-Iminodipropionic acid is a likely impurity. [*Beilstein RS*: **4 H 387**, **4 I 491**, **4 II 814**, **4 III 1222**, **4 IV 2481**; *R*: **4 H 385**, **4 I 491**, **4 II 812**, **4 III 1219**, **4 IV 2480**; *S*: **4 H 381**, **4 I 489**, **4 II 809**, **4 III 1208**, **4 IV 2480**.]

**β-Alanine** [107-95-9] **M 89.1**, **m 197-198°(dec)**, **205°(dec)**, **205.5°(dec)**, **207°(dec, rapid heating)**, **pK<sub>1</sub><sup>25</sup> 3.55 (3.60, CO<sub>2</sub>H)**, **pK<sub>2</sub><sup>25</sup> 10.24 (10.36, NH<sub>3</sub><sup>+</sup>)**. Crystallise β-alanine by dissolving it in a hot saturated aqueous solution, filtering, adding four volumes of absolute EtOH and cooling in an ice-bath. Recrystallise it in the same way and then finally, crystallise it from a warm saturated solution in 50% EtOH and adding four volumes of absolute EtOH with cooling in an ice-bath. The crystals are dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. It also crystallises from H<sub>2</sub>O, and sublimes at 170-180°/0.3mm. The *hydrochloride* [6057-90-5] forms plates with **m 123°**, its *methyl ester* has **b 69°/58mm**, the *methyl ester hydrochloride* [3196-73-4] has **m 107°** (from EtOH/Et<sub>2</sub>O), the *amide* has **m 41°** and the *amide hydrochloride* has **m 149°**. *N-Methyl β-alanine* crystallises from EtOH, the monohydrate forms plates with **m 99-100°** and its *hydrochloride* [2679-14-3] has **m 105°**. *N-Methyl β-alanine amide* [4874-17-3] **M 102.1** is a liquid with **d<sub>4</sub><sup>25</sup> 1.052**, **n<sub>D</sub><sup>20</sup> 1.458**, and is an antibacterial [Altamura et al. *J Med Chem* **38** 4244 1955]. [Donovan & Kegeles *J Am Chem Soc* **83** 255 1961, for pKa see Albert *Biochem J* **47** 531 1950, *Beilstein* **4 H 401**, **4 I 499**, **4 II 827**, **4 III 1258**, **4 IV 2526**.]

**S-Alaninol [S-2-aminopropan-1-ol]** [2749-11-3] **M 75.1**, **b 167-169°/760mm**, **d<sub>4</sub><sup>20</sup> 0.961**, **n<sub>D</sub><sup>20</sup> 1.456**, **[α]<sub>546</sub> +26.0° (c 2, EtOH)**, **pK<sub>25</sub> 9.43**. Purify it as for S-2-amino-3-methylbutan-1-ol below. [*Beilstein* **4 IV 1615**.]

**D-Allothreonine [2R,3R(-)-isomer]** [24830-94-2] **M 119.1**, **m 272-273°(dec)**, **276°(dec)**, **[α]<sub>D</sub><sup>25</sup> -9.1° (c 3.9, H<sub>2</sub>O)**, **pK<sub>1</sub><sup>25</sup> 2.11**, **pK<sub>2</sub><sup>25</sup> 9.10**. Recrystallise D-allothreonine from aqueous EtOH or 50% EtOH. [Elliot *J Chem Soc* **62** 1950, Birnbaum et al. *J Biol Chem* **194** 455 1952, IR: Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** 1961, *Beilstein* **4 IV 3170**.]

**RS-α-Allylglycine (2-aminopent-4-enoic acid)**. [7685-44-1] **M 115.1**, **m 250-255°(dec)**, **pK<sub>Est(1)</sub> ~2.3**, **pK<sub>Est(2)</sub> ~9.6**. Dissolve it in absolute EtOH and precipitate it with pyridine, then recrystallise it from aqueous EtOH [R<sub>F</sub> on paper in BuOH/EtOH/NH<sub>3</sub>/H<sub>2</sub>O (4:4:1:1) is 0.37]. The *hydrobromide* has **m 136-140°** (from EtOAc) and the *phenylureido* derivative has **m 159-161°**. [Schögl *Monatsh Chem* **89** 377 1958, *Beilstein* **4 IV 2852**.]

**Aminoacetic acid (Glycine)** [56-40-6] **M 75.1**, **m 262° (dec, goes brown at 226°, sublimes at 200°/0.1mm)**, **pK<sub>1</sub><sup>25</sup> 2.35**, **pK<sub>2</sub><sup>25</sup> 9.78**. Crystallise glycine from distilled water by dissolving at 90-95°, filtering, cooling to

about  $-5^{\circ}$ , and draining the crystals centrifugally. *Alternatively*, crystallise it from distilled water by addition of MeOH or EtOH (e.g. 50g dissolved in 100ml of warm water, and 400ml of MeOH is added). The crystals are washed with MeOH or EtOH, then with diethyl ether. Likely impurities are ammonium glycinate, iminodiacetic acid, nitrilotriacetic acid or/and ammonium chloride. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1955 1961, *Beilstein* 4 IV 2349.]

**$\alpha$ -Amino acids.** All the  $\alpha$ -amino acids with the “natural” configuration [*S* (L), except for cysteine which is *R*(L)] at the  $\alpha$ - carbon atom are available commercially in a very high state of purity. Many of the “non-natural”  $\alpha$ -amino acids with the [*R*(D)] configuration as well as racemic mixtures are also available, and generally none require further purification before use unless they are of “Technical Grade” or were stored for a very long period. The *R* or *S* enantiomers are optically active except for glycine which has two hydrogen atoms on the  $\alpha$ -carbon atom, but these are *pro*-chiral and enzymes or proteins do distinguish between them, e.g. serine hydroxymethyltransferase successfully replaces the *pro*- $\alpha$ -hydrogen atom of glycine with  $\text{CH}_2\text{OH}$  (from formaldehyde) to make *S*-serine. The twenty common natural  $\alpha$ -amino acids are: **amino acid**, three-letter abbreviation, **one-letter abbreviation**, pK (-COOH) and pK ( $-\text{NH}_3^+$ ): **Alanine**, Ala, A, 2.34, 9.69; **Arginine**, Arg, R, 2.17, 9.04; **Asparagine**, Asn, N, 2.01, 8.80; **Aspartic acid**, Asp, D, 1.89, 9.60; **Cysteine**, Cys, C, 1.96, 8.18; **Glutamine**, Gln, Q, 2.17, 9.13; **Glutamic acid**, Glu, E, 2.19, 9.67; **Glycine**, Gly, G, 2.34, 9.60; **Histidine**, His, H, 1.8, 9.17; **Isoleucine**, Ile, I, 2.35, 9.68; **Leucine**, Leu, L, 2.36, 9.60; **Lysine**, Lys, K, 2.18, 8.95; **Methionine**, Met, M, 2.28, 9.20; **Phenylalanine**, Phe, F, 1.83, 9.12; **Proline**, Pro, P, 1.99, 10.96; **Serine**, Ser, S, 2.21, 9.15; **Threonine**, Thr, T, 2.11, 9.62; **Tryptophan**, Trp, W, 2.38, 9.39; **Tyrosine**, Tyr, Y, 2.2, 9.11, **Valine**, Val, V, 2.32, 9.61, respectively. Technical grade amino acids can be purified on ion-exchange resins (e.g. Dowex 50W and eluting with a gradient of HCl or AcOH), and the purity is checked by TLC in two dimensions and stained with ninhydrin. (J.P. Greenstein & M. Winitz, *Chemistry of the Amino Acids* (3 Volumes), J. Wiley & Sons, NY, 1961; C. Cooper, N. Packer and K. Williams, *Amino Acid Analysis Protocols*, Humana Press, 2001, ISBN 0896036561). Recently codons for a further two amino acids have been discovered which are involved in ribosome-mediated protein synthesis giving proteins containing these amino acids. The amino acids are *R*(L)-selenocysteine [Stadtman *Ann Rev Biochem* 65 83 1996] and pyrrolysine [(4*R*, 5*R*)-4-substituted (with Me,  $\text{NH}_2$  or OH) pyrroline-5-carboxylic acid] [Krzycki & Chan et al. *Science* 296 1459 and 1462 2002.] They are, however, rare at present and only found in a few microorganisms.

***dl*- $\alpha$ -Amino adipic acid (hydrate) (2-amino hexane-1,6-dioic acid) [542-32-5] M 161.2, m 196-198 $^{\circ}$ , 204 $^{\circ}$ , 205-206 $^{\circ}$ , pK<sub>Est(1)</sub> ~2.0, pK<sub>Est(2)</sub> ~4.5, pK<sub>Est(3)</sub> ~9.8.** Crystallise the acid from  $\text{H}_2\text{O}$ . *Alternatively*, purify it by precipitating the Cu salt and decomposing the Cu salt suspended in  $\text{H}_2\text{O}$  by bubbling  $\text{H}_2\text{S}$ , filtering off the CuS, evaporating, and recrystallising the residue from  $\text{H}_2\text{O}$ . Note that prolonged refluxing of an aqueous solution converts the acid to the lactone: *piperid-2-one-6-carboxylic acid* which has m 177-178 $^{\circ}$ . [Linstead & Wang *J Chem Soc* 810, 811 1937, Waelkes et al. *J Am Chem Soc* 72 5760 1950, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p. 2408 1961, *Beilstein* 4 III 1555, 4 IV 3070.]

***N*-(*p*-Aminobenzoyl)-L-glutamic acid [4271-30-1] M 266.3, m 173 $^{\circ}$ , 174-175 $^{\circ}$  (L-form), [ $\alpha$ ]<sub>546</sub> -17.5 $^{\circ}$  (c 2, 0.1M HCl); 197 $^{\circ}$  (DL), pK<sub>1</sub><sup>25</sup> 2.61, pK<sub>2</sub><sup>25</sup> 3.76, pK<sub>3</sub><sup>25</sup> 4.83.** Crystallise the acid from  $\text{H}_2\text{O}$ . Also purify it by dissolving 2.7g in  $\text{H}_2\text{O}$  (130ml), adding aqueous NaOH to pH 5.5 and adding portionwise a solution of 0.5M  $\text{CuSO}_4$  to complete precipitation of the Cu salt. This salt is filtered off, suspended in  $\text{H}_2\text{O}$  and  $\text{H}_2\text{S}$  is bubbled through to precipitate CuS, filter, evaporate and recrystallise the residue from  $\text{H}_2\text{O}$ . It has  $\lambda_{\text{max}}$  ( $\text{H}_2\text{O}$ ) at 273nm. [Backer & Houtman *Rec Trav Chim Pays Bas* 70 738, 743 1951, *Beilstein* 14 IV 1153.]

***RS*-2-Aminobutyric acid [2835-81-6] M 103.1, m 287-288 $^{\circ}$ (dec), 303 $^{\circ}$ (dec), 303 $^{\circ}$ (dec, sealed tube), pK<sub>1</sub><sup>25</sup> 2.29, pK<sub>2</sub><sup>25</sup> 9.83.** Crystallise the acid from water. [Albertson & McKay *J Am Chem Soc* 81 505 1959, Perrin *J Chem Soc* 3125 1958, *Beilstein* 4 IV 2584.]

***S*-2-Aminobutyric acid (Butyrine) [1492-24-6] M 103.1, m 292 $^{\circ}$ (dec), [ $\alpha$ ]<sub>D</sub><sup>25</sup> + 20.6 $^{\circ}$  (c 2, 2.5N HCl), pK<sub>1</sub><sup>25</sup> 2.55, pK<sub>2</sub><sup>25</sup> 9.60.** Crystallise butyrine from aqueous EtOH, and the melting point depends on heating rate but has m 303 $^{\circ}$  in a sealed tube. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 2399 IR: 2401 1961, *Beilstein* 4 III 1294, 4 IV 2584.]

**RS-3-Aminobutyric acid** [2835-82-7] **M 103.1, m 193-194<sup>o</sup>, pK<sub>Est(1)</sub> ~3.5, pK<sub>Est(2)</sub> ~10.3.** Crystallise the acid from aqueous EtOH or MeOH/Et<sub>2</sub>O. Also crystallise it by heating a slightly diluted EtOH solution and adding Me<sub>2</sub>CO. It gives a purple spot with R<sub>F</sub> 0.89 on paper chromatography using 80% aqueous phenol (the α-amino acid has R<sub>F</sub> 0.74). [Zilka & Rivlin *J Org Chem* **23** 94 1958, Bruylants *Bull Soc Chim Belg* **32** 259 1923, Beilstein **4** IV 2595.]

**S-3-Aminobutyric acid** [3775-72-2] **M 103.1, m 212<sup>o</sup>, 210-212<sup>o</sup>, [α]<sub>D</sub><sup>20</sup> + 38.8<sup>o</sup> (c 0.5, H<sub>2</sub>O).** Purify the acid from absolute EtOH. It has also been crystallised from MeOH or MeOH/Et<sub>2</sub>O and dried in a vacuum. [Balenovic et al. *J Chem Soc* 3316 1952, Bruylants *Bull Soc Chim Belg* **32** 259 1923, Beilstein **4** IV 2595.]

**4-Aminobutyric acid (GABA)** [56-12-2] **M 103.1, m 202<sup>o</sup>(dec), 203<sup>o</sup>(dec), pK<sub>1</sub><sup>25</sup> 4.14, pK<sub>2</sub><sup>25</sup> 10.55.** Crystallise GABA from aqueous EtOH or MeOH/Et<sub>2</sub>O. Also crystallise it by dissolving it in the least volume of H<sub>2</sub>O and adding 5-7 volumes of absolute EtOH. [Sherman *Biochemical Preparations* **4** 91 1955, de Witt *Org Synth Coll Vol II* 25 1943, Beilstein **4** III 1316, **4** IV 2600.]

**1-Amino-1-cyclopentanecarboxylic acid (cycloleucine)** [52-52-8] **M 129.2, m 328-335<sup>o</sup>(dec), 328-329<sup>o</sup>, 330<sup>o</sup>(dec), pK<sub>1</sub><sup>20</sup> 2.4, pK<sub>2</sub><sup>20</sup> 10.3.** Any Cl<sup>-</sup> or other anions are removed by stirring with a strong cation exchange resin (Amberlite IR-120), filtering, and washing with distilled H<sub>2</sub>O until the filtrate is free from the anion. The resin is then stirred overnight with 6N NH<sub>4</sub>OH, filtered, the filtrate is decolourised (charcoal) and evaporated to dryness in a vacuum. The residue is recrystallised from H<sub>2</sub>O/EtOH. Also crystallise it from aqueous EtOH. The *hydrochloride* has **m 222-224<sup>o</sup>(dec)**. [Neelakantan & Hartung *J Org Chem* **23** 967 1958, Connors & Ross *J Chem Soc* 2119 1960, O'Donnell et al. *Synthesis* 127 1984, Beilstein **14** IV 974.]

**trans-(1RS,4RS)-4-Aminoethylcyclohexane-1-carboxylic acid (t-AMCHA, Tranexamic acid, Tranex, Transamin, Trasamlon, Ugurol, Frenolyse, Hexapromin, Abvitoff among other names)** [1197-18-8] **M 157.2, m 386-392<sup>o</sup> (dec), and cis-(1RS,4SR)-4-aminoethylcyclohexane-1-carboxylic acid** [1197-17-7] **M 157.2, m 236-238<sup>o</sup> (dec); cis- and trans- acids have the same pK<sub>1</sub><sup>25</sup> 4.51 (CO<sub>2</sub>H), pK<sub>2</sub><sup>25</sup> 10.72 (NH<sub>2</sub>).** A mixture of *cis*- and *trans*- acids (2:1) is obtained by catalytic reduction of 4-acetamidomethylbenzoic acid (3.9g, 20mmol, from acetylation of 4-aminomethylbenzoic acid see [56-91-7]) in aqueous NaOH (0.8g, 20mmol in 15ml of H<sub>2</sub>O) in the presence of Raney Ni (3ml of aqueous suspension) are shaken with H<sub>2</sub> in a bomb (100ml) at 170<sup>o</sup> and 82atm/cm<sup>2</sup> (1205psi). Reduction is complete after 2 hours, the catalyst is filtered off, the filtrate is acidified with 4N H<sub>2</sub>SO<sub>4</sub> and evaporated to dryness *in vacuo*; the residue is extracted with Me<sub>2</sub>CO, filtered and evaporated to dryness *in vacuo* to give crude 4-acetamidomethylcyclohexane-1-carboxylic acid (~3.9g). The acid is then refluxed with 20% HCl (20ml) for 3 hours (in an oil bath at 150<sup>o</sup>), evaporated to dryness *in vacuo*; the residue is dissolved in H<sub>2</sub>O (20ml) and freed from HCl by passage through a column of 'Amberlite IR-4B' (15ml, in OH<sup>-</sup> form) and eluted with H<sub>2</sub>O. Evaporation of the eluate *in vacuo* and crystallisation of the residue from aqueous Me<sub>2</sub>CO gives a 2:1 mixture of *cis*- and *trans*- **4-aminomethylcyclohexane-1-carboxylic acid** (2.24g, ~71%, **m ~232-236<sup>o</sup>, dec**). Both acids are symmetrical, i.e. mirror images are superimposable. The mixture (10g, 64mmol) is separated by refluxing with Cu(CO<sub>3</sub>)<sub>2</sub> (9.15g, 38mmol) in H<sub>2</sub>O (100ml) for 1 hour (turning deep blue in colour), cooled and the blue precipitate is filtered off, dissolved into 8% aqueous NH<sub>3</sub> (100ml), filtered from a little solid, and passed through a column of "Diaion SK#1 (NH<sub>4</sub><sup>+</sup> form)" and washed with H<sub>2</sub>O. The eluted solution is then passed through an "Amberlite IR-4B' (OH<sup>-</sup> form)" column and the effluent is evaporated to dryness. The residue (6.75g, **m ~221-223<sup>o</sup>, dec**.) is repeatedly recrystallised from H<sub>2</sub>O-Me<sub>2</sub>CO to give the pure *cis*-amino acid with **m 236-238<sup>o</sup>, dec**, and IR (KBr) peaks at ν<sub>max</sub> 2940, 2660, 1639 (1640), 1560 (1563), 1509 (1515), 1408 (1403), 1305 (1308), 930 and 904 cm<sup>-1</sup>, value in *italics* are from Meyer (see below). The combined filtrates from the recrystallisation are evaporated, the residue is dissolved in H<sub>2</sub>O and similarly de-ionised through the same columns. The final residue is recrystallised from H<sub>2</sub>O-Me<sub>2</sub>CO to give the pure *trans*-4-aminoethylcyclohexane-1-carboxylic acid (3.37g) with **m 286-292<sup>o</sup>, dec**, and IR (KBr) peaks at ν<sub>max</sub> 2940, 2610, 1637, 1535 (1528), 1383 (1381) (1325) and 920 cm<sup>-1</sup>, value in *italics* are from Meyer (see below). On TLC (Silica gel G, eluted with ascending *n*-PrOH/H<sub>2</sub>O, 65:35) the R<sub>F</sub> of the *cis*-isomer was always 1.2 times larger than that of the *trans*- isomer. [Note that the melting points of the isomers measured in the usual way vary somewhat, and by using a Du Pont Model 900 differential thermal analyzer the *cis*-acid had **m 252<sup>o</sup>**, and the *trans*-acid had **m 295-300<sup>o</sup>** as endotherms, Meyer *J Med Chem* **9** 641 1966]. The *cis*-hydrochloride has **m 195-197<sup>o</sup>, dec** (prisms from H<sub>2</sub>O-Me<sub>2</sub>CO), the *trans*-hydrochloride has **m 238-**

**241.5°**, **dec** (needles from H<sub>2</sub>O-Me<sub>2</sub>CO); the *cis*-hydrobromide has **m 205-208°**, **dec** (plates from Me<sub>2</sub>CO), the *trans*-hydrobromide has **m 227-229°**, **dec** (plates from H<sub>2</sub>O); the *cis*-*p*-toluenesulfonic acid salt has **m 177-178°** (plates from *n*-PrOH-Et<sub>2</sub>O), the *trans*-*p*-toluenesulfonic acid salt has **m 262-264°** (plates from H<sub>2</sub>O), the *cis*-HCl-AuCl<sub>3</sub> salt has **m 178-180°**, **dec** (yellow needles from H<sub>2</sub>O), the *trans*-HCl-AuCl<sub>3</sub> salt has **m 205-206°**, **dec** (yellow prisms from H<sub>2</sub>O), the *cis*-HCl-PtCl<sub>4</sub> salt has **m 233°**, **dec** (yellow needles from H<sub>2</sub>O), the *trans*-HCl-PtCl<sub>4</sub> salt has **m 254-255°**, **dec** (yellow-orange plates from H<sub>2</sub>O); the *cis*-*N*-acetamide has **m 189-190°** (prisms from EtOH), the *trans*-*N*-acetamide has **m 154-155°** (prisms from Me<sub>2</sub>CO); the *cis*-*N*-benzamide has **m 157-158°** (plates from EtOH/\*C<sub>6</sub>H<sub>6</sub>) and the *trans*-*N*-benzamide has **m 177-178°** (needles from EtOH/H<sub>2</sub>O).

The *cis*- and *trans*- acids are also readily separated from the mixture by recrystallisation of their *p*-toluenesulfonic acid salts which have very different solubility properties. **Isomerisation:** When a solution of the *cis*-acid (2g) in 0.5N NaOH (26ml) is heated in a silver vessel in an autoclave at 200° for 6 hours, and the cooled solution is passed through an "Amberlite IR-120" (NH<sub>4</sub><sup>+</sup> form) column and washed with H<sub>2</sub>O, and evaporation of the eluate *in vacuo* and four recrystallisations from aqueous MeOH, the *trans*-acid (0.8g, 40%), **m 384-390° dec**, is obtained [Naito et al. *Chem Pharm Bull Jpn* **16** 728 1968, Daiichi Seiyak Dutch Pat 6,414,942 1965, *Chem Abstr* **64** 3379 1966.] The width-at-half height of the peaks from the cyclohexane protons in the <sup>1</sup>H NMR spectra are larger from the *trans*-acid than the corresponding peaks from the *cis*-acid. The structure of the *cis*-acid was confirmed by conversion to a cyclic lactam on fusion. It has **m 104°** after recrystallisation from hexane followed by sublimation at 100°/2.5 x 10<sup>-2</sup>mm. [Note that this melting point was recorded at the point where, under crossed Nichol prisms, birefringence was lost. The crystal form, however, was only slowly lost thereafter over a wide range of temperatures]. The lactam has IR (KBr) peaks at  $\nu_{\max}$  1661 (amide I), 1421, 1325 and 1205 cm<sup>-1</sup> [Meyer *J Med Chem* **9** 641 1966].

These amino acids are haemostatic with antiplasmin activity. The *trans*-acid has potent antiplasmin activity (inhibiting the fibrinolytic enzyme system), being 50 times more active than the *cis*-acid, and 5-7 times more potent than  $\epsilon$ -aminocaproic acid [Naito et al. *Chem Pharm Bull Jpn* **16** 357 and 728 1968.] The antifibrinolytic activity is due to blocking of the lysine binding sites of plasminogen. *t*-AMCHA has been used as a lysine analogue to characterise binding sites in plasminogens. [Hoover et al. *Biochemistry* **32** 10936 1993, Marshall et al. *Biochemistry* **33** 3599 1994.]

**4-Amino hippuric acid** (*N*-*p*-aminobenzoylglycine) [61-78-9] **M 194.2**, **m 198-199°**, **200-202°**, **pK<sub>Est(1)</sub> ~1.7(NH<sub>2</sub>)**, **pK<sub>Est(2)</sub> ~3.4(CO<sub>2</sub>H)**. Crystallise the acid from H<sub>2</sub>O. It is soluble in organic solvents. [Cohen & McGilvery *J Biol Chem* **169** 119 1945, **171** 121 1947, Meunzen et al. *J Biol Chem* **26** 469 1926, *Beilstein* **14** III 1069, **14** IV 1152.]

***dl*-4-Amino-3-hydroxybutyric acid** [924-49-2] **M 119.1**, **m 218°(dec)**, **225°(dec)**, **pK<sub>25</sub> ~3.80(CO<sub>2</sub>H)**, **pK<sub>Est(2)</sub> ~9.3**. Crystallise the acid from H<sub>2</sub>O or aqueous EtOH. Recrystallise it by dissolving it in H<sub>2</sub>O and adding MeOH or EtOH. It is not very soluble in CHCl<sub>3</sub> or EtOAc. [Renaud & Seebach *Synthesis* 424 1986, *Beilstein* **4** IV 3187.]

***R* (L-)-4-Amino-3-hydroxybutyric acid (GABOB)** [352-21-6] **M 119.1**, **m 212°(dec)**, **213-214°(dec)**, **216-217°(dec)**, **[ $\alpha$ ]<sub>D</sub><sup>20</sup> -20.5° (c 1.75, H<sub>2</sub>O)**. Purify GABOB through a Dowex 50Wx8 resin, eluting with 1.3N NH<sub>4</sub>OH, evaporating and crystallising the residue by dissolving it in H<sub>2</sub>O and adding EtOH. It is an anticonvulsant. [Renaud & Seebach *Synthesis* 424 1986, Takano et al. *Tetrahedron Lett* **29** 795 1988, *Beilstein* **4** IV 3187.]

**$\alpha$ -Aminoisobutyric acid (2-amino-2-methylpropionic acid)** [62-57-7] **M 103.1**, **m sublimes at 280-281°**, **335° (sealed tube)**, **pK<sub>25</sub> 2.36**, **pK<sub>25</sub> 10.21**. Crystallise the acid from aqueous EtOH and dry it at 110°. [Zelinski & Stadnikoff *Chem Ber* **39** 1726 1906, *Beilstein* **4** IV 2616.]

***RS*- $\beta$ -Aminoisobutyric acid ( $\alpha$ -methyl- $\beta$ -alanine)** [10569-72-9] **M 103.1**, **m 176-178°**, **178-180°**, **181-182°**, ***R*-(-)- isomer [144-90-1] m 183°**, **[ $\alpha$ ]<sub>D</sub><sup>20</sup> -21° (c 0.43, H<sub>2</sub>O)**, **pK<sub>Est(1)</sub> ~3.7**, **pK<sub>Est(2)</sub> ~10.2**. *RS*- $\beta$ -Aminoisobutyric acid forms colourless prisms by crystallisation from hot H<sub>2</sub>O which are powdered and dried *in vacuo*. The purity is checked by paper chromatography (Whatman 1) using ninhydrin spray to visualise the amino acid; R<sub>F</sub> values in 95% MeOH and *n*-PrOH/5N HCOOH (8:2) are 0.36 and 0.50 respectively. [Kupiecki & Coon *Biochemical Preparations* **7** 20 1960, Pollack *J Am Chem Soc* **65** 1335 1943.] The *R*-enantiomer, isolated from

iris bulbs or human urine, crystallises from H<sub>2</sub>O and sublimes *in vacuo* [Asen et al. *J Biol Chem* **234** 343 1959]. The *RS*-hydrochloride crystallises from EtOH/Et<sub>2</sub>O with *m* 128-129° (also 130°) [Böhme et al. *Chem Ber* **92** 1258, 1260, 1261 1959]. [Beilstein **4** III 1330.]

**5-Aminolaevulinic acid hydrochloride (ALA-HCl,  $\delta$ -aminolaevulinic acid HCl)** [5451-09-2] *M* 167.6, *m* 148°(dec), 150-151°(dec), 156-158°(dec), *pK*<sub>1</sub><sup>22</sup> 4.05, *pK*<sub>2</sub><sup>22</sup> 8.90. Dry ALA-HCl in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> overnight, then crystallise it by dissolving it in cold EtOH and adding dry Et<sub>2</sub>O. Also crystallise it by dissolving in the minimum volume of MeOH, and placing in a desiccator containing dry Et<sub>2</sub>O (clamp the desiccator). During several days the Et<sub>2</sub>O slowly distils into the MeOH causing the hydrochloride to separate as long needles. Filter them off and dry them in a Fischer pistol. [Wynn & Corwin *J Org Chem* **15** 203, 207 1950, Neuberger & Scott *J Chem Soc* 1820, 1924 1954, Beilstein **4** IV 3265.]

**4-Aminomethylbenzoic acid ( $\alpha$ -amino-*p*-toluic acid)** [56-91-7] *M* 151.2, *m* 273-274°, 294-295°, >300°, 345°, 347-350° (sealed tube), *pK*<sub>1</sub><sup>20</sup> 3.59 (CO<sub>2</sub>H), *pK*<sub>2</sub><sup>20</sup> 9.64 (NH<sub>2</sub>). This acid has been prepared in two different ways from *p*-cyanobenzoic acid [Levine & Sedlecky *J Org Chem* **24** 115 1959, Albert & Magrath *J Chem Soc* 678 1944], and by reduction of *p*-carboxybenzaldehyde oxime [Nair & Baugh *J Org Chem* **38** 2189 1973]. A mixture of *p*-cyanobenzoic acid (14g, 619-65-8), Raney cobalt (2g, W-6 or W-7), 28% aqueous NH<sub>3</sub> (40ml) and H<sub>2</sub>O (150ml) are shaken in a Parr hydrogenator at 3 atm (initial pressure) and 25° for ~3 hours when the theoretical volume of H<sub>2</sub> is absorbed. The catalyst is filtered off, the filtrate is boiled to remove NH<sub>3</sub> and the solid that separated is collected, and recrystallised from H<sub>2</sub>O (charcoal) to give the amino acid (*m* 347-350°) in 80% yield. Alternatively, the oxime of *p*-carboxybenzaldehyde (1g, 619-66-9) in 95% EtOH (100ml) containing 5% Pd/C (100mg) is shaken with H<sub>2</sub> at 30 psi for 18 hours at 25°. The catalyst is filtered off and washed with hot glacial AcOH (2 x 20ml); the combined washings and filtrate are evaporated to dryness and the residue is recrystallised from H<sub>2</sub>O (charcoal) to give the white crystalline amino acid (850mg, *m* 294-295°). The melting point appears to vary with heating rate. UV has  $\lambda_{\text{max}}$  at 234nm (H<sub>2</sub>O), and the <sup>1</sup>H NMR [TFA] has  $\delta$  at 4.15 (q, *J* = 6Hz, benzylic H), 7.3 (d, *J* = 8Hz, two protons adjacent to H<sub>2</sub>NCH<sub>2</sub>) and 7.9 (d, *J* = 8Hz, two protons adjacent to CO<sub>2</sub>H). The *N*-acetyl derivative crystallises from a large volume of xylene and has *m* 199-200° (lit: Levine & Sedlecky state *m* 199-120°). [For *pK*<sub>a</sub> see Goldacre *Nature* **154** 796 1944, Beilstein **14** H 487, **14** III 1212, **14** IV 1362.]

***S*-2-Amino-3-methyl-1-butanol (*S*-valinol)** [2026-48-4] *M* 103.2, *m* 31-32°, *b* 88°/11mm, *d* 0.92, [ $\alpha$ ]<sub>546</sub> + 16.5° (*c* 6.32, *l* = 2 H<sub>2</sub>O), [ $\alpha$ ]<sub>D</sub> + 15.6° (EtOH), *pK*<sub>Est</sub> ~10.4. Purify *S*-valinol by vacuum distillation using a short Vigreux column. Alternatively, it is purified by steam distillation. The steam distillate is acidified with HCl; the aqueous layer is collected and evaporated. The residue is dissolved in butan-1-ol, filtered and dry Et<sub>2</sub>O added to crystallise the hydrochloride salt (*hygroscopic*), *m* 113°. The free base can be obtained by suspending the salt in Et<sub>2</sub>O and adding small volumes of saturated aqueous K<sub>2</sub>CO<sub>3</sub> until effervescence is complete and the mixture is distinctly alkaline. At this stage the aqueous layer should appear as a white sludge. The mixture is heated to boiling and refluxed for 30 minutes (more Et<sub>2</sub>O is added if necessary). The Et<sub>2</sub>O layer is decanted off from the white sludge, the sludge is extracted twice with Et<sub>2</sub>O (by boiling for a few minutes), the combined organic layers are dried (KOH pellets), evaporated and the residue is distilled in a vacuum. [Nagao et al. *J Org Chem* **55** 1148 1990, Beilstein **4** III 805.]

***S*-(+)-3-Aminopentanoic acid** [14389-77-6] and ***R*-(-)-3-aminopentanoic acid** [131347-76-7] *M* 115.1, *m* (175°), 185°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> (+) and (-) 43° (*c* 0.5, H<sub>2</sub>O), *pK*<sub>1</sub><sup>25</sup> 3.54, *pK*<sub>2</sub><sup>25</sup> 10.25. Crystallise the amino acid from EtOH/Et<sub>2</sub>O. [Beilstein **4** II 843, **4** III 1342, **4** IV 2635.]

**$\alpha$ -Aminothiophene-2-acetic acid [2-(2-thienyl)glycine]** [*R*(+) 65058-23-3, *S*(-) 4052-59-9, (-)- 43189-45-3, *RS*( $\pm$ ) 21124-40-3] *M* 57.2, *m* 236-237° (*R*), 235-236° (*S*), 208-210°, 223-224° (dec)(*RS*), [ $\alpha$ ]<sub>D</sub><sup>20</sup> (+) and (-) 84° (*c* 1, 1% aqueous HCl), [ $\alpha$ ]<sub>D</sub><sup>25</sup> (+) and (-) 71° (*c* 1 H<sub>2</sub>O), *pK*<sub>Est(1)</sub> ~ 1.5, *pK*<sub>Est(2)</sub> ~ 8.0. Recrystallise 2-(2-thienyl)glycine by dissolving it in H<sub>2</sub>O (1g in 3 ml), adjusting the pH to 5.5 with aqueous NH<sub>3</sub>, diluting with MeOH (20 ml), stirring, adjusting the pH to 5.5 and cooling to 0°. Also recrystallise it from small volumes of H<sub>2</sub>O. [*R*-isomer: Nishimura et al. *Nippon Kagaku Zasshi* **82** 1688 1961, *S*-isomer: Johnson & Panetta *Chem*

*Abstr* **63** 14869 1965, Johnson & Hardcastle *Chem Abstr* **66** 10930 1967, *RS*-isomer: LiBassi et al. *Gazz Chim Ital* **107** 253 1977.] The ( $\pm$ ) *N*-acetyl derivative has *m* 191° (from H<sub>2</sub>O) [Schouteenten et al. *Bull Soc Chim Fr* II 248, II 252 1978].

**5-Amino-*n*-valeric acid (5-aminopentanoic acid)** [660-88-8] *M* 117.2, *m* 157-158°, *pK*<sub>1</sub><sup>25</sup> 4.25, *pK*<sub>2</sub><sup>25</sup> 10.66. Crystallise it from H<sub>2</sub>O/EtOH. When heated above its melting point, it is converted to 2-piperidone with *m* 200°. [Wood & Colver *J Am Chem Soc* **67** 654 1945, *Beilstein* **4** IV 2636.]

**5-Amino-*n*-valeric acid hydrochloride** [627-95-2] *M* 153.6, *m* 92-94°, 103-104°. Crystallise the salt from CHCl<sub>3</sub>. Otherwise dissolve it in EtOH and add 2 volumes of Et<sub>2</sub>O and chill. [Schniepp & Marvel *J Am Chem Soc* **57** 1557 1935, Woods & Colver *J Am Chem Soc* **67** 654 1945, *Beilstein* **4** III 1343, **4** IV 2636.]

**Angiotensinogen (from human blood serum or porcine plasma)** [64315-16-8] *M* ~1762. A tetradecapeptide renin substrate which is purified by chromatography on Blue Sepharose, Phenyl-Sepharose, hydroxylapatite and immobilised 5-hydroxytryptamine [Campbell et al. *Biochem J* **243** 121 1987]. [*Beilstein* **25** III/IV 4390 for *angiotensin*.]

**Anserine [N, $\beta$ -alanyl-1-methyl-*S*-histidine]** [584-85-0] *M* 240.3, *m* 238-239°, [ $\alpha$ ]<sub>D</sub><sup>23</sup> +12.3° (c 5, H<sub>2</sub>O), *pK*<sub>1</sub><sup>25</sup> 2.64, *pK*<sub>2</sub><sup>25</sup> 7.04, *pK*<sub>3</sub><sup>25</sup> 9.49. Crystallise anserine from aqueous EtOH. It is *hygroscopic* and is best stored as the nitrate salt (see below). Purify it by shaking the nitrate salt with Dowex 3 (x4 free base) and washing with H<sub>2</sub>O, evaporating the filtrate and removing H<sub>2</sub>O by 3 distillations with 10ml of propan-2-ol. Dissolve the crystals in MeOH and add H<sub>2</sub>O dropwise until one phase is obtained and cool. Dry the crystals at 60° over P<sub>2</sub>O<sub>5</sub> in a vacuum. The *picrate* has *m* 145° (from H<sub>2</sub>O). [Rinderknecht et al. *J Org Chem* **29** 1968 1964, *Beilstein* **25** II 408, **25** IV 4383.]

***S*-Anserine nitrate** [5937-77-9] *M* 303.3, *m* 225°(dec), 226-228°(dec), [ $\alpha$ ]<sub>D</sub><sup>30</sup> +12.2°. Likely impurities are 1-methylimidazole-5-alanine and histidine. Crystallise the nitrate from aqueous MeOH or EtOH (needles). Also dissolve ~20g in 25ml of MeOH, add 2-propanol (150-200ml) and store the mixture at 5° overnight to give shiny needles. Recrystallise it by heating 12g of the nitrate in MeOH (300ml) and adding H<sub>2</sub>O (50-60ml) until one phase is obtained and refrigerating overnight. Filter and dry it at 60°/P<sub>2</sub>O<sub>5</sub> in a vacuum. [Rinderknecht et al. *J Org Chem* **29** 1968 1964, Behrens & duVigneaud *J Biol Chem* **120** 517 1937.]

***S*-Arginine** [74-79-3] *M* 174.2, *m* 205°(dec, anhydrous), 207°(dec, 2 H<sub>2</sub>O), [ $\alpha$ ]<sub>D</sub> +26.5° (c 5, in 5M HCl), [ $\alpha$ ]<sub>546</sub> +32° (c 5, in 5M HCl), *pK*<sub>1</sub><sup>25</sup> 2.18, *pK*<sub>2</sub><sup>25</sup> 9.36, *pK*<sub>3</sub><sup>25</sup> 11.5. *S*-Arginine crystallises from H<sub>2</sub>O as the *dihydrate* and as plates from EtOH. It also crystallises from 66% EtOH. Its solubility in H<sub>2</sub>O is 15% at 21°. Its isoelectric point is at pH 10.76. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1841 1961, *Beilstein* **4** IV 817.]

***S*-Arginine hydrochloride** [1119-34-2] *M* 210.7, *m* 217°(dec), 222°(dec), [ $\alpha$ ]<sub>D</sub><sup>30</sup> +26.9° (c 6, M HCl). A likely impurity is ornithine. Crystallise the salt from H<sub>2</sub>O at pH 5-7, by adding EtOH to 80% (v/v). [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1841 1961, *Beilstein* **4** IV 2649.]

***S*-Argininosuccinic acid** [2387-71-5] *M* 290.3, [ $\alpha$ ]<sub>D</sub><sup>23</sup> +16.4° (H<sub>2</sub>O). A likely impurity is fumaric acid. In neutral or alkaline solution it readily undergoes ring closure to the "anhydride" (see below). Crystallise it from water by adding 1.5 volumes of EtOH. The barium salt is stable at 0-5° if dry. [Westfall *Biochem J* **77** 135 1960, Ratner & Kunkemueller *Biochemistry* **5** 1821 1966.]

***S*-Argininosuccinic anhydride** [28643-94-9] *M* 272.3, [ $\alpha$ ]<sub>D</sub><sup>23</sup> -10° (H<sub>2</sub>O for anhydride formed at neutral pH). Crystallise the anhydride from H<sub>2</sub>O by adding two volumes of EtOH. An isomeric anhydride is formed if the free acid is allowed to stand at acid pH. In solution, the mixture of anhydride and free acid is formed [see above entry, Ratner & Kunkemueller *Biochemistry* **5** 1821 1966, Kowalsky & Ratner *Biochem J* **8** 899 1969].

***S*-Asparagine** [70-47-3] *M* 150.1, *m* 234-235° (monohydrate) [5794-13-8] [ $\alpha$ ]<sub>D</sub> +32.6° (0.1M HCl), *pK*<sub>1</sub><sup>25</sup>

**1.98, pK<sub>2</sub><sup>25</sup> 8.84.** Likely impurities are aspartic acid and tyrosine. Crystallise it from H<sub>2</sub>O or aqueous EtOH. It slowly effloresces in dry air. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1856 1961, Beilstein 4 IV 3005.]

**Aspartic acid** M 133.1, m 338-339° (*RS*, [617-45-8]), m 271° (*S*, requires heating in a sealed tube [56-84-8]) [α]<sub>D</sub><sup>25</sup> +25.4° (3M HCl), pK<sub>1</sub><sup>25</sup> 1.99, pK<sub>2</sub><sup>25</sup> 3.90. Likely impurities are glutamic acid, cystine and asparagine. Crystallise the acid from water by adding 4 volumes of EtOH and dry it at 110°. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 1856 1961, Beilstein 4 IV 2998, 3000.]

**L-Aspartic acid β-methyl ester hydrochloride** [16856-13-6] M 183.6, m 194°, pK<sup>25</sup> 8.62. Recrystallise it from MeOH by adding anhydrous Et<sub>2</sub>O [Bach et al. *Biochemical Preparations* 13 20 1971].

**DL-Aspartic acid dimethyl ester hydrochloride** [14358-33-9] M 197.7, 116-117°. Crystallise it from absolute MeOH. [Kovach et al. *J Am Chem Soc* 107 7360 1985.] The *diethyl ester* has pK<sup>25</sup> 6.4.

**Azaserine** [115-02-6] M 173.1, m 146-162°(dec), [α]<sub>D</sub><sup>27.5</sup> -0.5° (c 8.5, H<sub>2</sub>O, pH 5.2), pK<sub>Est(1)</sub> ~4.53, pK<sub>Est(2)</sub> ~5.40. Crystallise azaserine from 90% EtOH. Also dissolve it in H<sub>2</sub>O, filter it through Supercel and add EtOH to give azaserine as pale yellow crystals. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 1 pp 75-76 1961, Curphey & David *J Org Chem* 43 4666 1978, Beilstein 4 IV 3124.]

**Benzoyl glycine (hippuric acid)** [495-69-2] M 179.2, m 188°, pK<sup>20</sup> 3.81, 3.59. pK<sup>40</sup> 3.59. Crystallise the acid from boiling H<sub>2</sub>O. Dry it over P<sub>2</sub>O<sub>5</sub>. Also purify it by dissolving 135-140g in 2L of boiling H<sub>2</sub>O, filtering through a steam-heated funnel and allowing to crystallise at ~20° (yield 115-122g first crop, m 186-187°). [Ingersoll & Babcock *Org Synth Coll Vol II* 328 1943, Beilstein 9 225, I 100.]

**N-Benzyloxycarbonylglycyl-L-alaninamide** [17331-79-2] M 279.3, m dec >200°. Recrystallise the dipeptide derivative from EtOH/Et<sub>2</sub>O.

**N-Benzyloxycarbonyl-N'-methyl-L-alaninamide** [33628-84-1] M 236.3, m dec >200°. Recrystallise the amide from EtOAc.

**Betaine (1-carboxymethyl-N,N,N-trimethylammonium zwitterion)** [107-43-7 (*anhydrous*), 590-47-6, 17146-86-0 (*monohydrate*)] M 117.1, m 294-294°(dec) (*monohydrate?*) 301-305°(dec) (*anhydrous*), ~319°(dec), pK<sup>25</sup> 1.83. Crystallise betaine from aqueous EtOH or EtOH/Et<sub>2</sub>O. The *monohydrate* loses H<sub>2</sub>O above 100°. Betaine undergoes internal alkylation to methyl dimethylaminoacetate above its melting point. It is also prepared by treating the hydrochloride (below) with silver oxide and recrystallising from EtOH/Et<sub>2</sub>O. [Edsall *J Am Chem Soc* 66 1767 1943, Leifer & Lippincott *J Am Chem Soc* 79 5098 1957, for pK see Grob et al. *Chem and Ind (London)* 1222 1955, Beilstein 4 III 1127, 4 IV 2369.]

**Betaine hydrochloride** [590-46-5] M 153.6, m 227-228°(dec), 232°(dec), 246-247°(dec). Recrystallise the salt from EtOH. Its solubility at 25° is 65% in H<sub>2</sub>O, and 5% in EtOH. [Edsall *J Am Chem Soc* 66 1767 1943, Kuhn & Ruelius *Chem Ber* 83 420 1950, Beilstein 4 III 1127, IV 2369.]

**Bis-N-tert-butylloxycarbonyl-L-cystine**, [10389-65-8] M 440.5, m 144.5-145°, [α]<sub>D</sub><sup>20</sup> -133.2° (c 1, MeOH), pK<sub>Est</sub> ~2.9. Crystallise the cystine derivative from EtOAc by adding hexane [Ferraro *Biochemical Preparations* 13 39 1971].

**Bombesin (2-L-glutamine-3,6-L-asparaginealytesin, a tetradecapeptide)** [31362-50-2] M 1619.9. Purify Bombesin by gel filtration on a small column of Sephadex G-10 and elute with 0.01 M AcOH. This procedure removes lower molecular weight contaminants which are retarded on the column. The procedure should be repeated twice, and the material should now be homogeneous on electrophoresis, and on chromatography it gives a single active spot which is negative to ninhydrin but positive to Cl<sub>2</sub> and iodoplatinate reagents. R<sub>F</sub> on paper chromatography (*n*-BuOH/pyridine/AcOH/H<sub>2</sub>O :: 37.5: 25:7.5:30) is 0.55 for Bombesin and 0.65 for

Alytin. [Bernardi et al. *Experientia* **Part 1** **27** 166 1971, Anastasi et al. **Part 2** **27** 873 1971.] The hydrochloride has  $m$  185°(dec) (from EtOH)  $[\alpha]_{\text{D}}^{25} -20.6^{\circ}$  [c 0.65, Me<sub>2</sub>NCHO/(Me<sub>2</sub>N)<sub>3</sub>PO (8:2)]. [For the stimulation of inositol phosphate see Lloyd et al. *Biochem J* **260** 813 1989.]

**Bradykinin** [ArgProProGlyPheSerProPheArg] [5979-11-3]  $M_r$  1,240.4. Purify Bradykinin by ion-exchange chromatography on CMC (*O*-carboxymethyl cellulose) and partition chromatography on Sephadex G-25. The purity is checked by paper chromatography using BuOH/AcOH/H<sub>2</sub>O (4:1:5) as eluent. [Park et al. *Can J Biochem* **56** 92 1978, ORD and CD: Bodanszky et al. *Experientia* **26** 948 1970, activity: Regoli & Barabé *Pharmacol Rev* **32** 1 1980, *Beilstein* **22** III/IV 91.]

**S-Canavanine** [2-amino-4-(guanidinoxy)butyric acid] [543-38-4]  $M$  176.2,  $m$  184°,  $[\alpha]_{\text{D}}^{17} +19.4^{\circ}$  (c 2, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{20} +7.9^{\circ}$  (c 3.2, EtOH),  $pK_1^{25}$  2.43,  $pK_2^{25}$  6.60,  $pK_3^{25}$  9.25. Crystallise *S*-canavanine from absolute EtOH or aqueous EtOH. [Tomiyama *J Biol Chem* **111** 48 1935 gave  $pK_1^{25}$  9.25 (COOH),  $pK_2^{25}$  7.4 (guanidinium),  $pK_3^{25}$  11.5 (NH<sub>4</sub><sup>+</sup>), Gulland & Morris *J Chem Soc* 763 1935, (±) Frankel et al. *J Chem Soc* 3127 1963, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2622-2628 1961, *Beilstein* **4** III 1636, **4** IV 3188.]

**S-Canavanine sulfate** (from jackbean, *O*-guanidino-*L*-homoserine) [2219-31-0]  $M$  274.3,  $m$  160-165°(dec), 172°(dec),  $[\alpha]_{\text{D}}^{18.5} +19.8^{\circ}$  (c 7, H<sub>2</sub>O),  $pK_1^{25}$  7.40 (CO<sub>2</sub>H),  $pK_2^{25}$  9.25 (α-NH<sub>2</sub>),  $pK_3^{25}$  11.5 (guanidinoxy). Recrystallise the sulfate by dissolving (~1g) in H<sub>2</sub>O (10ml), and adding with stirring 0.5 to 1.0 volumes of 95% EtOH whereby crystals separate. These are collected, washed with Me<sub>2</sub>CO/EtOH (1:1) and dried over P<sub>2</sub>O<sub>5</sub> in a vacuum. [Hunt & Thompson *Biochemical Preparations* **13** 416 1971, Feacon & Bell *Biochem J* **59** 221 1955, *Beilstein* **4** III 1636, **4** IV 3188.]

***N*-Carbamoylglycine** (hydantoic acid, *N*-carboxymethylurea, ureidoacetic acid) [462-60-2]  $M$  118.0,  $m$  169-170°(dec),  $m$  173-175°(dec), 178-180°(dec),  $pK_a$  (H<sub>2</sub>O) 3.91 (5°), 3.80 (20.3°), 3.87 (32.1°) and 3.89 (50°). Hydantoic acid is prepared by reaction of potassium cyanate (KNCO) with glycine in H<sub>2</sub>O, and is recrystallised from H<sub>2</sub>O or EtOH. [West *J Biol Chem* **34** 187 1918, Dakin *J Chem Soc* **107** 434 1915, King *J Am Chem Soc* **78** 6020 1956, Inouye & Watanabe *J Chem Soc, Perkin Trans I* 1911 1977.] It has UV  $\lambda_{\text{max}}$  (ε) at 215 (2800) and 240 (165)nm (neutral species: EtOH); 215 (2600) and 240 (105)nm (cation: EtOH + H<sup>+</sup>); 215 (2400) and 240 (80)nm (anion: EtOH + OH<sup>-</sup>) [Crombie & Hooper *J Chem Soc* 3010 1955]. The ethyl ester [6293-20-5]  $M$  146.0, crystallises from H<sub>2</sub>O as needles with  $m$  135°; and the amide [3530-79-8]  $M$  117.1, crystallises from H<sub>2</sub>O as prisms with  $m$  204° (180° has also been reported). [*Beilstein* **4** H 359, **4** I 477, **4** II 792, **4** III 1163, **4** IV 2411.]

**Carnitine** (α-hydroxy-β-*N,N,N*-trimethylaminopropionic acid) [*R*(+) 541-14-0, *S*(-) 541-15-1, *RS* 461-06-3]  $M$  161.2,  $m$  *R* or *S* isomer 197-198°(dec), 210-212°(dec), *RS* isomer 195-197°,  $[\alpha]_{\text{D}}^{20}$  (+) and (-) 36° (c 10, H<sub>2</sub>O),  $pK^{25}$  3.6. The *S*(*L*) isomer is levocarnitine, Vitamin B<sub>7</sub>. The *R* or *S* isomers crystallise from EtOH/Me<sub>2</sub>CO (hygroscopic). The *R* or *S* hydrochlorides crystallise from hot EtOH or EtOH/Et<sub>2</sub>O and have  $m$  142°(dec). The *RS*-isomer crystallises from hot EtOH (hygroscopic). The *RS* hydrochloride crystallises in needles from hot EtOH and has  $m$  196°(dec). [(±) Mazzetti & Lemmon *J Org Chem* **22** 228 1957, *Beilstein* **4** H 513, **4** I 548, **4** II 937-8, **4** III 1632-5, **4** IV 3185.]

**L-Carnosine** (β-alanyl-*L*-histidine) [305-84-0]  $M$  226.2,  $m$  258-260°(dec), 260°(capillary tube), 262°(dec),  $[\alpha]_{\text{D}}^{25} +20.5^{\circ}$  (c 1.5, H<sub>2</sub>O),  $pK_1^{25}$  2.64,  $pK_2^{25}$  6.83,  $pK_3^{25}$  9.51. Likely impurities are histidine and β-alanine. Crystallise *L*-carnosine from water by adding EtOH in excess. Recrystallise it from aqueous EtOH by slow addition of EtOH to a strong aqueous solution of the dipeptide. Its solubility in H<sub>2</sub>O is 33.3% at 25°. [Vinick & Jung *J Org Chem* **48** 392 1983, Turner *J Am Chem Soc* **75** 2388 1953, Sifford & du Vigneaud *J Biol Chem* **108** 753 1935, *Beilstein* **25** H 516, **25** I 717, **25** II 408.]

**S-Citrulline** (2-amino-5-ureidopentanoic acid) [372-75-8]  $M$  175.2,  $m$  222°,  $[\alpha]_{\text{D}}^{20} +24.2^{\circ}$  (in 5M HCl),  $pK^{25}$  2.43,  $pK_2^{25}$  9.41. Likely impurities are arginine and ornithine. Crystallise *S*-citrulline from water by adding 5 volumes of EtOH. Also crystallise it from water by addition of MeOH. [Ellenbogen *J Am Chem Soc*

74 5198 1952, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2491-2494 1961, Beilstein 4 IV 2647.]

**Corticotropin** [92307-52-3] **polypeptide**  $M_r$  ~4697. The extract is purified by ion-exchange on CM-cellulose, desalted, evaporated and lyophilised. Then separate from impurities by gel filtration through Sephadex G-50. [Lande et al. *Biochemical Preparations* 13 45 1971, Esch et al. *Biochem Biophys Res Commun* 122 899 1984].

**Creatine (N-guanidino-N-methylglycine)** [6020-87-7 (*monohydrate*), 57-00-1 (*anhydrous*)] **M 131.1 (anhydrous), 149.1 (hydrate) m 303<sup>o</sup>, pK<sub>1</sub><sup>25</sup> 2.63, pK<sub>2</sub><sup>25</sup> 14.3.** Likely impurities are creatinine and other guanidino compounds. It crystallises from the minimum volume of boiling H<sub>2</sub>O as the *monohydrate*. The *hydrate* is also obtained by dissolving in H<sub>2</sub>O and adding Me<sub>2</sub>CO. Drying under vacuum over P<sub>2</sub>O<sub>5</sub> or drying at 100<sup>o</sup> gives the *anhydrous* base. The *anhydrous* base can be obtained also by dissolving the hydrate in H<sub>2</sub>O, seeding with the anhydrous base and cooling in ice. A **m of 258-268<sup>o</sup>(dec)** was reported. The *picrate* crystallises from 17 parts of H<sub>2</sub>O with **m of 218-220<sup>o</sup>(dec)**. [King *J Chem Soc* 2377 1930, Hoffmann et al. *J Am Chem Soc* 58 1730 1936, Mendel & Hodgkin *Acta Cryst* 7 443 1954, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 2750 1961, Beilstein 4 III 1170, 4 IV 2425.]

**Creatine phosphate di Na, 4H<sub>2</sub>O salt (phosphocreatine)** [922-32-7] **M 327.1, pK<sub>1</sub><sup>27</sup> 2.7, pK<sub>2</sub><sup>27</sup> 4.58, pK<sub>3</sub><sup>27</sup> ~12.** To 3-4g of the salt in H<sub>2</sub>O (220ml) is added 4 volumes of EtOH with thorough stirring and allowed to stand at 20<sup>o</sup> for 12 hours (this temperature is critical as crystals do not readily form at 23<sup>o</sup> or 25<sup>o</sup>). The salt first appears as oily droplets which slowly settle and crystallise. After 12 hours the supernatant is clear. Stirring and scratching the flask containing the filtrate brings out additional crystals (0.3-1g) if the salt is kept at 20<sup>o</sup> for 12 hours. Filter it off at room temperature, wash with 3 x 5ml of ice-cold 90% EtOH, then 5ml of absolute EtOH and dry it in a vacuum desiccator (Drierite or CaCl<sub>2</sub>) for 16-30 hours. The *hexahydrate* (plates) is converted to the *tetrahydrate* salt (needles) in a vacuum at -10<sup>o</sup>. [Ennor & Stocken *Biochemical Preparations* 5 9 1957, *Biochem J* 43 190 1958, Beilstein 4 III 1170, 4 IV 2425.]

**Creatinine [2-amino-1-methylimidazolidin-4-one, 2-imino-1-methyl-4(3H)-oxoimidazolidine]** [60-27-5] **M 113.1, m ~305<sup>o</sup>(dec), pK<sub>1</sub><sup>25</sup> 4.80, pK<sub>2</sub><sup>25</sup> 9.2.** Likely impurities are creatine and ammonium chloride. Dissolve it in dilute HCl, then neutralise with ammonia. Recrystallise it from H<sub>2</sub>O by adding excess of Me<sub>2</sub>CO. The *picrate* crystallises from 23 volumes of boiling H<sub>2</sub>O and has **m 220-221<sup>o</sup>(dec)**. [King *J Chem Soc* 2377 1930, Beilstein 25 III/IV 2543.]

**D-(R-natural) and L-(S-non-natural) Cycloserine (2-amino-3-isoxazolidone)** [R- 68-41-7 and S- 339-72-0] **M 102.1, m 145-150<sup>o</sup> (dec), 154-155<sup>o</sup>, 155-156<sup>o</sup> (dec), 156<sup>o</sup> (dec), [α]<sub>D</sub><sup>25</sup> (+) and (-) 137<sup>o</sup> (c 5, 2N NaOH), pK<sub>1</sub><sup>10</sup> 4.5, pK<sub>2</sub><sup>10</sup> 7.74, pK<sub>3</sub><sup>25</sup> 4.50, pK<sub>4</sub><sup>25</sup> 7.43, pK<sub>1</sub><sup>50</sup> 4.44, pK<sub>2</sub><sup>50</sup> 7.20.** Purify cycloserine by recrystallisation from aqueous EtOH or MeOH or aqueous NH<sub>3</sub>/EtOH or isoPrOH. Also recrystallise it from aqueous ammoniacal solution at pH 10.5 (100mg/ml) by diluting with 5 volumes of isopropanol and then adjusting to pH 6 with acetic acid. An aqueous solution, buffered to pH 10 with Na<sub>2</sub>CO<sub>3</sub>, can be stored in a refrigerator for 1 week without decomposition. Its UV has λ<sub>max</sub> at 226nm (*A*<sub>1cm</sub><sup>1%</sup> 4.02). The *tartrate salt* has **m 165-166<sup>o</sup> (dec), 166-168<sup>o</sup> (dec), and [α]<sub>D</sub><sup>24</sup> -41<sup>o</sup> (c 0.7, H<sub>2</sub>O)**. [Stammer et al. *J Am Chem Soc* 79 3236 1959, UV: Kuehl *J Am Chem Soc* 77 2344 1955, Beilstein 27 III/IV 5549.]

**Cystamine dihydrochloride [2,2'-diaminodiethylene disulfide dihydrochloride, 2,3'-dithio-bis(ethylamine) dihydrochloride]** [56-17-7] **M 225.2, m 219-220<sup>o</sup>(dec), pK<sub>1</sub><sup>30</sup> 8.82, pK<sub>2</sub><sup>30</sup> 9.58.** Recrystallise the salt by dissolving in EtOH containing a few drops of dry EtOH/HCl, filtering and adding dry Et<sub>2</sub>O. The solid is dried in a vacuum and stored in a dry and dark atmosphere. It has been recrystallised from EtOH (solubility: 1g in 60ml of boiling EtOH) or MeOH (plates). The *free base* has **b 90-100<sup>o</sup>/0.001mm, 106-108<sup>o</sup>/5mm and 135-136<sup>o</sup>/760mm, d<sub>4</sub><sup>20</sup> 1.1559, n<sub>D</sub><sup>20</sup> 1.5720.** [Verly & Koch *Biochem J* 58 663 1954, Gonick et al. *J Am Chem Soc* 76 4671 1954, Jackson & Block *J Biol Chem* 113 137 1936.] The *dihydrobromide* has **m 238-239<sup>o</sup> (from EtOH/Et<sub>2</sub>O)** [Viscontini *Helv Chim Acta* 36 835 1953]. [Beilstein 4 H 287, 4 IV 1578.]

**S,S-(L,L)-Cystathionine** (*S*-2-amino-2-carboxyethyl-L-homocysteine, L-2-amino-4[(2-amino-2-carboxyethyl)thio]butyric acid) [56-88-2] **M 222.3**, **m** >300°, **dec** at 312° with darkening at 270°,  $[\alpha]_D^{20} +23.9^\circ$  (c 1, **M HCl**). *S,S*-Cystathionine is purified by converting it to the *HCl* salt in 20% *HCl* and carefully basifying with aqueous  $\text{NH}_3$  until separation is complete. Filter it off and dry it in a vacuum. It forms prisms from  $\text{H}_2\text{O}$ . The *dibenzoyl* derivative has **m** 229° (from *EtOH*). [IR: Greenstein & Winitz *Chemistry of the Amino Acids* (*J Wiley*) **Vol 3** p2690 1961 and Tallan et al. *J Biol Chem* **230** 707 1958, Synthesis: du Vigneaud et al. *J Biol Chem* **143** 59 1942, Anslow et al. *J Biol Chem* **166** 39 1946.] [Prepn: Weiss & Stekol *J Am Chem Soc* **73** 2497 1951; see also du Vigneaud et al. *J Biol Chem* **143** 60 1942, Biological synthesis: Greenberg *Methods Enzymol* **5** 943 1962, *Beilstein* **4** IV 3197.]

**Cysteamine** (2-aminoethanethiol, 2-mercaptoethylamine) [60-23-1] **M 77.2**, **m** 97-98.5°, 98-99°, 99-100°, **pK<sub>1</sub><sup>0</sup> 9.15**, **pK<sub>2</sub><sup>0</sup> 11.93**, **pK<sub>1</sub><sup>30</sup> 8.42**, **pK<sub>2</sub><sup>30</sup> 10.83**. It is soluble in  $\text{H}_2\text{O}$  giving an alkaline reaction, and it has a disagreeable odour. A likely impurity is the disulfide *cystamine* which is not soluble in alkaline solution. Under a  $\text{N}_2$  atmosphere dissolve it in *EtOH*, evaporate to dryness and wash the white residue with dry petroleum ether, then sublime it at 0.1mm and store it under  $\text{N}_2$  at 0-10° in the dark. Its  $\text{HgCl}_2$  (2:3) complex has **m** 181-182° (from  $\text{H}_2\text{O}$ ), and its *picrate* has **m** 125-126°. [Mills & Bogert *J Am Chem Soc* **57** 2328 1935, **62** 1173 1940, Baddiley & Thain *J Chem Soc* 800 1952, Shirley *Preparation of Organic Intermediates* (*J. Wiley*) **Vol 3** 189 1951, Barkowski & Hedberg *J Am Chem Soc* **109** 6989 1987, *Beilstein* **4** IV 1570.]

**Cysteamine hydrochloride** [156-57-0] **M 113.6**, **m** 70.2-70.7°, 70-72°. Purify the salt by recrystallisation from *EtOH*. It is freely soluble in  $\text{H}_2\text{O}$  and should be stored in a dry atmosphere. [Mills & Bogert *J Am Chem Soc* **62** 1177 1940.] The *picrate* has **m** 125-126°; see previous entry for *free base*. [*Beilstein* **4** IV 1570.]

(±)-**Cysteic acid** (3-sulfoalanine, 1-amino-3-sulfopropionic acid) [13100-82-8, 3024-83-7] **M 169.2**, **m** 260°(dec). Likely impurities are cystine and oxides of cysteine. Crystallise the acid from water by adding 2 volumes of *EtOH*. It crystallises from  $\text{H}_2\text{O}$  as the *monohydrate*. When recrystallised from aqueous *MeOH* it has **m** 264-266°, and the *anhydrous* acid has **m** ~260°(dec). [Chapeville & Formageot *Biochim Biophys Acta* **26** 538 1957, Gortner & Hoffman *J Biol Chem* **72** 435 1927, Greenstein & Winitz *The Chemistry of the Amino Acids* *J. Wiley*, **Vol 3** p1908 1961.]

**R(L)-Cysteic acid (H<sub>2</sub>O)** [23537-25-9] **M 187.2**, **m** 275-280°(dec), 289°,  $[\alpha]_D^{20} +8.66^\circ$  (c 7.4,  $\text{H}_2\text{O}$ , **pH 1**) and +1.54° ( $\text{H}_2\text{O}$ , **pH 13**), **pK<sub>1</sub><sup>25</sup> 1.9** ( $\text{SO}_3\text{H}$ ), **pK<sub>2</sub><sup>25</sup> 8.7** ( $\text{CO}_2\text{H}$ ), **pK<sub>3</sub><sup>25</sup> 12.7** ( $\text{NH}_2$ ). Likely impurities are cystine and oxides of cysteine. Crystallise it from water by adding 2 volumes of *EtOH*. When recrystallised from aqueous *MeOH* it has **m** 264-266°, and the *anhydrous* acid has **m** ~260°(dec). [Chapeville & Formageot *Biochim Biophys Acta* **26** 538 1957, Riordan & Giese *Methods Enzymol* **47** 31 1977, *Beilstein* **4** IV 3296.]

**D-(S)- and L-(R)- Cysteine** (*S*- and *R*-2-amino-3-mercaptopropionic acid) [*S*(+) 921-01-7, *R*(-) 52-90-4] **M 121.2**, **m** 230°, 240° (dec),  $[\alpha]_D^{20} +$  and  $-7.6^\circ$  (c 2, **M HCl**), + and  $-10.1^\circ$  (c 2,  $\text{H}_2\text{O}$ , **pH 10**), **pK<sub>1</sub><sup>25</sup> 1.92** ( $\text{CO}_2$ ), **pK<sub>2</sub><sup>25</sup> 8.35** ( $\text{NH}_2$ ), **pK<sub>3</sub><sup>25</sup> 10.46** (**SH**). Purify it by recrystallisation from  $\text{H}_2\text{O}$  (free from metal ions) and dry it in a vacuum. It is soluble in  $\text{H}_2\text{O}$ , *EtOH*,  $\text{Me}_2\text{CO}$ , *EtOAc*, *AcOH*,  $\text{C}_6\text{H}_6$  and  $\text{CS}_2$ . Acidic solutions can be stored under  $\text{N}_2$  for a few days without deterioration. [For synthesis and spectra see Greenstein & Winitz *Chemistry of the Amino Acids* (*J. Wiley*) **Vol 3** p1879 1961, *Beilstein* **4** III 1618, **4** IV 3144.]

**L-Cysteine hydrochloride (H<sub>2</sub>O)** [52-89-1 (*anhydrous*), 7048-04-6 (*monohydrate*)] **M 175.6**, **m** 175-178° (dec),  $[\alpha]_D^{25} +6.53^\circ$  (**5M HCl**). Likely impurities are cystine and tyrosine. Crystallise the salt from *MeOH* by adding diethyl ether, or from hot 20% *HCl*. Dry it under vacuum over  $\text{P}_2\text{O}_5$ . *Hygroscopic*. [*Beilstein* **4** III 1580, 1600.]

(±)-**Cysteine hydrochloride** [10318-18-0 (*anhydrous*), 116797-51-5 (*monohydrate*)] **M 157.6**, **m** 140-141.5° (dec), **pK<sub>2</sub><sup>20</sup> 8.36** ( $\text{NH}_2$ ), **pK<sub>2</sub><sup>20</sup> 10.28** (**SH**). Crystallise the salt from hot 20% *HCl* and dry it under vacuum over  $\text{P}_2\text{O}_5$ . It also crystallises from *EtOH* with **m** 175° (hydrate?). When crystallised from absolute *EtOH* or *EtOH/Et<sub>2</sub>O*, it has **m** 140-141.5° (*anhydrous*?). [Rurner & Voitle *J Am Chem Soc* **72** 628 1950, Albert *Biochem J* **50** 690 1953, *Beilstein* **4** IV 3145.]

**L-Cystine** [56-89-3] M 240.3,  $[\alpha]_D^{25} -229^{\circ}$  (c 0.92 in M HCl),  $\text{pK}_1^{25} 1.04$  (1.65),  $\text{pK}_2^{25} 2.05$  (2.76),  $\text{pK}_3^{25} 8.00$  (7.85),  $\text{pK}_4^{25} 10.25$  (8.7, 9.85). Cystine disulfoxide impurity is removed by treating an aqueous suspension with  $\text{H}_2\text{S}$ . The cystine is filtered off, washed with distilled water and dried at  $100^{\circ}$  under a vacuum over  $\text{P}_2\text{O}_5$ . Crystallise it by dissolving in 1.5M HCl, then adjusting to neutral pH with ammonia. Likely impurities are D-cystine, meso-cystine and tyrosine. Also purify it by dissolving it in 10%  $\text{NH}_3$  and adding gradually dilute AcOH until the point of precipitation and cooling slowly [Dughton & Harrison *Acta Cryst* **12** 396, 402 1959.] Alternatively dissolve it in 6N  $\text{NH}_4\text{OH}$  and evaporate it at room temperature for crystallisation to occur. [Chaney & Steinrauf *Acta Cryst* **30** 711 1974, *Beilstein* **4** IV 3155.]

**meso-2,6-Diaminopimelic acid** [583-93-7] M 190.2, m 313-315 $^{\circ}$ (dec)  $\text{pK}_1^{25} 1.04$  (1.65, 1.8),  $\text{pK}_2^{25} 2.05$  (2.2, 2.76),  $\text{pK}_3^{25} 8.00$  (7.85, 8.8),  $\text{pK}_4^{25} 10.25$  (8.7, 9.85, 9.9), pI ~5.5. Crystallise the acid from  $\text{H}_2\text{O}$  or aqueous EtOH. Also purify it by dissolving it in hot  $\text{H}_2\text{O}$  and adding 5 volumes of EtOH, filter after 12 hours at  $-10^{\circ}$ . The acid has been recrystallised from 35% aqueous EtOH [Wade et al. *J Am Chem Soc* **79** 648, 651 1957]. [*Beilstein* **4** IV 3081.]

**L(S)-2,3-Diaminopropionic acid monohydrochloride (3-amino-L-alanine hydrochloride)** [1482-97-9] M 140.6, m 132-133 $^{\circ}$ (dec), 237 $^{\circ}$ (dec),  $[\alpha]_D^{25} +26.1^{\circ}$  (c 5.8, M HCl),  $\text{pK}_1^{25} 1.30$ ,  $\text{pK}_2^{25} 6.79$ ,  $\text{pK}_3^{25} 9.51$ . It forms needles from  $\text{H}_2\text{O}$  and can be recrystallised from aqueous EtOH. [Gmelin et al. *Z Physiol Chem.* **314** 28 1959, IR: Koegel et al. *J Am Chem Soc* **77** 5708 1977, *Beilstein* **4** IV 2501.]

**meso-2,3-Diaminosuccinic acid** [23220-52-2] M 148.1, m 305-306 $^{\circ}$ (dec, and sublimes),  $\text{pK}_{\text{Est}(1)} \sim 3.6$ ,  $\text{pK}_{\text{Est}(2)} \sim 9.8$ . Crystallise the acid from water. Also, dissolve it in dilute NaOH and add AcOH to pH 5-6 and allow it to crystallise (m 304 $^{\circ}$  dec). *Alternatively*, dissolve the acid in aqueous  $\text{NH}_3$  and boil; when the  $\text{NH}_3$  has evaporated, the acid separates, filter it off and dry it at room temperature in a vacuum. In another procedure 1g of acid is dissolved in 10ml of conc HCl + 15ml of  $\text{H}_2\text{O}$  at  $80^{\circ}$ , filter immediately, dilute with 20ml of  $\text{H}_2\text{O}$  and allow to stand for 24 hours. When the *monohydrochloride* (0.7g, m 175-156 $^{\circ}$  dec) crystallises out, filter and dry it. It has also been purified by dissolving it in the minimum volume of 10% HCl, filtering, and diluting with 5 volumes of  $\text{H}_2\text{O}$  when the crystals separate slowly on standing. The acid is filtered off after 24 hours and dried (m 306-306 $^{\circ}$  dec). Similar procedures were used for the *dl-isomer*. [Kenner *J Org Chem* **13** 28 1948, McKennis & Yard *J Org Chem* **23** 980 1958, *Beilstein* **4** III 1528, **4** IV 3025.]

**6-Diazo-5-oxo-L-norleucine** [157-03-9] M 171.2, m 140-150 $^{\circ}$ (dec), 145-155 $^{\circ}$ (dec),  $[\alpha]_D^{20} +21^{\circ}$  (c 5, EtOH),  $\text{pK}_1 2.1$ ,  $\text{pK}_2 8.95$ . Crystallise it from EtOH,  $\text{H}_2\text{O}/\text{EtOH}$ , MeOH, 95% aqueous MeOH or  $\text{H}_2\text{O}/\text{Me}_2\text{CO}$ . [DeWald & Moor *J Am Chem Soc* **80** 3944 1958, Dion et al. *J Am Chem Soc* **78** 3075 1956, *Beilstein* **4** IV 3278.]

**Diglycyl glycine (triglycine)** [556-33-2] M 189.2, m 246 $^{\circ}$ (dec),  $\text{pK}_1^{25} 3.30$ ,  $\text{pK}_2^{25} 7.96$ . Crystallise triglycine from  $\text{H}_2\text{O}$  or  $\text{H}_2\text{O}/\text{EtOH}$  and dry it at  $110^{\circ}$ . [Yakel & Hughes *Acta Cryst* **5** 847 1952, Lenel et al. *Acta Cryst* **3** 313 1952, Holley & Holley *J Am Chem Soc* **74** 3072 1952, *Beilstein* **4** III 1198, **4** IV 2469.]

**N,N-Di-(2-hydroxyethyl)glycine (BICINE, N,N-bis-(2-hydroxyethyl)glycine)** [150-25-4] M 163.2, m 193 $^{\circ}$ (dec), 193-195 $^{\circ}$ (dec),  $\text{pK}_1^{25} 2.50$ ,  $\text{pK}_2^{25} 8.11$ . Dissolve bicine in a small volume of hot water and precipitate it with EtOH, twice. Repeat once more but treat the aqueous solution with charcoal and filter before adding EtOH. Also crystallise it from concentrated aqueous solutions. [Torn & Kolthoff *J Am Chem Soc* **77** 2061 1955, Chaberek et al. *J Am Chem Soc* **75** 2185 1953, *Beilstein* **4** IV 2390.]

**3-(3,4-Dihydroxyphenyl)-L-alanine (DOPA, LEVODOPA, EUODOPA)** [59-92-7, 5796-17-8] M 197.2, m 275 $^{\circ}$ (dec), 267-268 $^{\circ}$ (dec), 284-286 $^{\circ}$ (dec), ~295 $^{\circ}$ (dec),  $[\alpha]_D^{15} -13.1^{\circ}$  (c 5.12, N HCl),  $\text{pK}_1^{25} 2.32$  ( $\text{CO}_2\text{H}$ ),  $\text{pK}_2^{25} 8.72$  ( $\text{NH}_2$ ),  $\text{pK}_3^{25} 9.96$  (OH),  $\text{pK}_4^{25} 11.79$  (OH). Likely impurities are vanillin, hippuric acid, 3-methoxytyrosine and 3-aminotyrosine. DOPA recrystallises from large volumes of  $\text{H}_2\text{O}$  forming colourless white needles; its solubility in  $\text{H}_2\text{O}$  is 0.165%, but it is insoluble in EtOH,  $^*\text{C}_6\text{H}_6$ ,  $\text{CHCl}_3$ , and EtOAc. Also crystallise it by dissolving it in dilute HCl and adding dilute ammonia to give pH 5, under  $\text{N}_2$ . *Alternatively*, crystallise it from dilute aqueous EtOH. It is rapidly oxidised in air when moist, and darkens, particularly in

alkaline solution. Dry it *in vacuo* at 70° in the dark, and store it in a dark container preferably under N<sub>2</sub>. It has  $\lambda_{\max}$  at 220.5nm (log  $\epsilon$  3.79) and 280nm (log  $\epsilon$  3.42) in 0.001N HCl. [Yamada et al. *Chem Pharm Bull Jpn* **10** 693 1962, Bretschneider et al. *Helv Chim Acta* **56** 2857 1973, NMR: Jardetzky & Jardetzky *J Biol Chem* **233** 383 1958, *Beilstein* **4** IV 2492, 2493.]

**3-(3,4-Dihydroxyphenyl)-2-methyl-L-alanine** [methyl dopa, **2-amino-3-(3,4-dihydroxy-phenyl)-2-methylpropionic acid**] [555-30-6, 41372-08-1 (*sesquihydrate*)] **M 238.2, m >300°, 300-301°(dec), pK<sub>1</sub><sup>25</sup> 2.2, pK<sub>2</sub><sup>25</sup> 9.2, pK<sub>3</sub><sup>25</sup> 10.6, pK<sub>4</sub><sup>25</sup> 12.0.** Recrystallise methyl dopa from H<sub>2</sub>O. [Reinhold et al. *J Org Chem* **33** 1209 1968.] The *L*-isomer forms a *sesquihydrate* from H<sub>2</sub>O **m 302-304° (dec)**, and the *anhydrous* crystals are *hygroscopic*,  $[\alpha]_{\text{D}}^{23}$  -4.0° (c 1, 0.1N HCl),  $[\alpha]_{546}$  +154.5° (c 5, CuSO<sub>4</sub> solution). It has  $\lambda_{\max}$  at 281nm ( $\epsilon$  2780). Its solubility in H<sub>2</sub>O at 25° is ~10mg/ml and the pH of an aqueous solution is ~5.0. It is insoluble in most organic solvents. [Stein et al. *J Am Chem Soc* **77** 700 1955, *Beilstein* **4** IV 2505.]

**3,5-Diiodo-L-thyronine (3,5-diiodo-4-[4-hydroxyphenoxy]-1-phenylalanine)** [1041-01-6] **M 525.1, m 255°(dec), 255-257°(dec),  $[\alpha]_{\text{D}}^{22}$  +26° [2N HCl-EtOH (1:2)], pK<sub>1</sub><sup>20</sup> 3.25, pK<sub>2</sub><sup>20</sup> 5.32, pK<sub>3</sub><sup>20</sup> 9.48.** Recrystallise it from EtOH. [Chambers et al. *J Chem Soc* 3424 1949, *Beilstein* **14** III 1565, **14** IV 2372.]

**3,5-Diiodo-L-tyrosine dihydrate** [300-39-0] **M 469.0, m 199-210°, 202°(dec), 204°(dec),  $[\alpha]_{\text{D}}^{20}$  +2.89° (c 4.9, 4% HCl), pK<sub>1</sub><sup>25</sup> 2.12, pK<sub>2</sub><sup>25</sup> 6.48, pK<sub>3</sub><sup>25</sup> 7.82.** It forms crystals from H<sub>2</sub>O [solubility (g/L): 0.204 at 0°, 1.86 at 50°, 5.6 at 75° and 17.0 at 100°]. Also recrystallise it from 50% or 70% EtOH. When boiled in EtOH the crystals swell, and on further boiling a gelatinous precipitate is formed [Harrington *Biochem J* **22** 1434 1928, Jurd *J Am Chem Soc* **77** 5747 1955]. It also crystallises from cold dilute ammonia on adding acetic acid to pH 6. [*Beilstein* **14** IV 2370.]

**dl-4-Dimethylamino-2,2-diphenylvaleramide (Dimevamide, Aminopentamide)** [60-46-8] **M 296.4, m 183-184°, pK<sub>Est</sub> ~9.8.** Crystallise dimevamide from aqueous EtOH. The *hydrochloride* forms leaflets from EtOH/Et<sub>2</sub>O with **m 190-191°** and is deliquescent. The *picrate* has **m 210-211°**. It is an antispasmodic. [Moffett et al. *J Am Chem Soc* **79** 4451, 4457 1957, Beckett *J Pharm Pharmacol* **8** 848 1956, *Beilstein* **14** III 1363, **14** IV 1865.]

**(-)-L-4-Dimethylamino-2,2-diphenylvaleramide** [6078-64-4] **M 296.4, m 134.5-135.5°, 136.5-137.5°,  $[\alpha]_{\text{D}}^{20}$  -112° (c 0.87, EtOH), -84.1° (c 0.9, 0.04N HCl), pK<sub>Est(1)}</sub> 8.3.** Crystallise the amide from petroleum ether, EtOH or as needles from aqueous EtOH. It is an analgesic. [Beckett et al. *J Chem Soc* 3076 1957.]

**N,N-Dimethylglycinehydrazide hydrochloride** [539-64-0] **M 153.6, m 181°.** Crystallise the salt by adding EtOH to a concentrated aqueous solution. [Viscontini & Meier *Helv Chim Acta* **33** 1773 1950, *Beilstein* **4** III 1127, **4** IV 2368.]

**Djenkolic acid (S,S'-methylene-bis-L-cysteine)** [498-59-9] **M 254.3, m 300-350°(dec),  $[\alpha]_{\text{D}}^{20}$  -65° (c 2, HCl) [See pK of S-methyl-L-cysteine].** Crystallise djenkolic acid from a large volume of water (solubility is 0.5g%). [du Vigneaud & Patterson *J Biol Chem* **114** 533 1936, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2682, 2687 1961, *Beilstein* **4** III 1591.] The *N,N'*-dibenzoyl derivative crystallises with 1H<sub>2</sub>O from aqueous EtOH with **m 87.5-89°** [*Beilstein* **9** III 1171.]

**S-Ethionine** [13073-35-3] **M 163.2, m 282°(dec),  $[\alpha]_{\text{D}}^{25}$  +23.7° (in 5M HCl), pK<sup>25</sup> 9.02 (for RS).** Likely impurities are *N*-acetyl-(*R* and *S*)-ethionine, *S*-methionine, and *R*-ethionine. Crystallise it from water by adding 4 volumes of EtOH or 85% aqueous EtOH. It sublimes at 196-216°/0.3mm with 99.1% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. [Weiss & Stekol *J Am Chem Soc* **73** 2399 1951, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2658, 2659 1961, *Beilstein* **4** IV 3194.]

**Ethylene N,N'-bis[(*o*-hydroxyphenyl)glycine]** [1170-02-1] **M 360.4, m 249°(dec), pK<sub>Est(1)}</sub> ~1.8, pK<sub>Est(2)}</sub> ~4.8, pK<sub>Est(3)}</sub> ~9.0.** Purify it by extensive Soxhlet extraction with acetone. [Bonadies & Carrano *J Am Chem*

*Soc* 108 4088 1986].

**2-Fluorophenylalanine** [*R*(+) 97731-02-7, *S*(-) 19883-78-4] **M 183.2, m 226-232°**, 231-234°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> (+) and (-) 15° (c 2, H<sub>2</sub>O pH 5.5), pK<sub>1</sub><sup>24</sup> 2.12, pK<sub>2</sub><sup>24</sup> 9.01. Recrystallise 2-fluorophenylalanine from aqueous EtOH. The *hydrochloride* has m 226-231°(dec), and the *N-acetyl* derivative has m 147-149° (aqueous EtOH). [Bennett & Nieman *J Am Chem Soc* 72 1800 1950, *Beilstein* 14 III 1268.]

**4-Fluorophenylalanine** [*R*(+) 18125-46-7, *S*(-) 1132-68-9] **M 183.2, m 227-232°**, [ $\alpha$ ]<sub>D</sub><sup>25</sup> + and -24° (c 2, H<sub>2</sub>O), pK<sub>1</sub><sup>24</sup> 2.13, pK<sub>2</sub><sup>24</sup> 9.05. It is recrystallised from aqueous EtOH. The (*R*)-*N-acetyl* derivative has m 142-145°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -38.6° (c 8, EtOH). [Bennett & Nieman *J Am Chem Soc* 72 1800 1950, *Beilstein* 14 III 1268.]

**L-5-Fluorotryptophan monohydrate** [16626-02-1; 154-08-5] **M 240.2, m reported for the L-enantiomer 158-163°(dec)**, (±)-isomer >250°(dec), 264-265°(also 238-239° dec reported), [ $\alpha$ ]<sub>D</sub><sup>20</sup> +5.5° (c 1, 0.1N HCl), pK<sub>Est(1)</sub>~2.5 (CO<sub>2</sub>H), pK<sub>Est(2)</sub>~9.4 (NH<sub>2</sub>), pK<sub>Est(3)</sub>~16 (indole-NH). Recrystallise it from EtOH, aqueous EtOH or AcOH. Also purify it by passage through a Dowex AG1x2 (acetate form) column and recrystallise the L-enantiomer (from enzymic enrichment) from H<sub>2</sub>O/EtOH, m 158-163°(dec), [ $\alpha$ ]<sub>D</sub><sup>23</sup> -8.3° (c 2.5, N NaOH). [Coy et al. *Biochemistry* 13 3550 1974, *Beilstein* 22/14 V 116.]

**L-Glutamic acid** [56-86-0] **M 147.1, m 224-225°(dec)**, [ $\alpha$ ]<sub>D</sub><sup>25</sup> +31.4° (c 5, 5M HCl), pK<sub>1</sub><sup>20</sup> 2.06, pK<sub>2</sub><sup>20</sup> 4.35, pK<sub>3</sub><sup>20</sup> 9.85. Crystallise L-glutamic acid from H<sub>2</sub>O acidified to pH 3.2 by adding 4 volumes of EtOH, and drying at 110°. Likely impurities are aspartic acid and cysteine. It sublimes at 170-175°/10mm. It melts at 160° with cyclisation to L-pyrrolidone carboxylic acid. [Dunn & Brophy *J Biol Chem* 99 224 1958, Parikh et al. *J Am Chem Soc* 80 957 1958, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 1929-1952 1961, *Beilstein* 4 III 1530, 4 IV 3028.]

**L-Glutamic acid- $\gamma$ -benzyl ester** [1676-73-9] **M 237.3, m 179-181°**, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +19.3° (c 1, AcOH), pK<sub>1</sub><sup>25</sup> 2.17, pK<sub>2</sub><sup>25</sup> 9.00. Recrystallise the ester from H<sub>2</sub>O and store it at 0°. [Estrin *Biochemical Preparations* 13 25 1971, *Beilstein* 6 IV 2538.]

**L-Glutamine** [56-85-9] **M 146.2, m 184-185°**, 187°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> +31.8° (M HCl), pK<sub>1</sub><sup>25</sup> 2.17, pK<sub>2</sub><sup>25</sup> 9.13. Likely impurities are glutamic acid, ammonium pyroglutamate, tyrosine, asparagine, isoglutamine, arginine. Crystallise it from water or aqueous EtOH. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 1929-1925 1961, *Beilstein* 4 IV 3038.]

**L-Glutathione (reduced form,  $\gamma$ -L-glutamyl-L-cysteinyl-glycine)** [70-18-8] **M 307.3, m 188-190°(dec)**, 195°(dec), [ $\alpha$ ]<sub>D</sub><sup>20</sup> -20.1° (c 1, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 2.12 (CO<sub>2</sub>H), pK<sub>2</sub><sup>25</sup> 3.59 (CO<sub>2</sub>H), pK<sub>3</sub><sup>25</sup> 8.75 (NH<sub>2</sub>), pK<sub>4</sub><sup>25</sup> 9.65 (10.0, SH). Crystallise L-glutathione from 50% aqueous EtOH, dry it in a vacuum and store it below 5°. Alternatively, recrystallise it from aqueous EtOH under N<sub>2</sub>, and store it dry in a sealed container below 4°. It is soluble in H<sub>2</sub>O. [Weygand & Geiger *Chem Ber* 90 634 1957, Martin & Edsall *Bull Soc Chim Fr* 40 1763 1958, du Vigneaud & Miller *Biochemical Preparations* 2 87 1952, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 2 p 1523 1961, *Beilstein* 4 IV 3165.]

**L-Glutathione (oxidised)** [27025-41-8] **M 612.6, m 175-195°**, 195°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -98° (c 2, H<sub>2</sub>O), pK<sub>1</sub> 3.15, pK<sub>2</sub> 4.03, pK<sub>3</sub> 8.75. Purify it by recrystallisation from 50% aqueous EtOH. Its solubility in H<sub>2</sub>O is 5%. Store it at 4°. [Li et al. *J Am Chem Soc* 76 225 1954, Berse et al. *Can J Chem* 37 1733 1959, *Beilstein* 4 IV 3168.]

**Glycinamide hydrochloride** [1668-10-6] **M 110.5, m 186-189°**, 203-205°, 207-208°, pK<sub>1</sub><sup>25</sup> -6.10, pK<sub>2</sub><sup>25</sup> -1.78, pK<sub>3</sub><sup>25</sup> 7.95. Crystallise the salt from EtOH, EtOH/H<sub>2</sub>O or MeOH. [Karmas & Spoerri *J Am Chem Soc* 74 1580 1952, *Beilstein* 4 IV 2358.]

**Glycine anhydride (2,5-diketopiperazine)** [106-57-0] **M 114.1, m 309-310°**, 311-312°(dec), ~315°(dec), pK<sub>1</sub> -4.45, pK<sub>2</sub> -2.16 (pK<sub>2</sub> -1.94 in AcOH). Recrystallise glycine anhydride from H<sub>2</sub>O (plates) and it can be

sublimed (slowly) at 260° or at 140-170°/0.5mm. The *dihydrochloride* has **m** 129-130° and is prepared by dissolving it in conc HCl and adding EtOH to crystallisation point; dry it in a vacuum. The *bis-1-naphthylurethane* has **m** 232°(dec), and the *diperchlorate* has **m** 117° (*hygroscopic*). [MS: Johnstone *J Chem Soc, Perkin Trans 1* 1297 1975, NMR: Blaha & Samek *Col Czech Chem Commun* 32 3780 1967, Sauborn *J Phys Chem* 36 179 1932, Corey *J Am Chem Soc* 60 1599 1938, *Beilstein* 24 IV 1070.]

**Glycine ethyl ester hydrochloride** [623-33-6] **M 136.9, m 145-146°, pK<sup>25</sup> 7.69**. Crystallise it from absolute EtOH or EtOH/Et<sub>2</sub>O. [Marvel *Org Synth Coll Vol II* 310 1943, *Beilstein* 4 II 780, 4 III 3 75.]

**Glycine tert-butyl ester (glycine 1,1-methylethyl ester)** [6456-742] **M 131.1, 65-67°/20mm, d<sub>D</sub><sup>20</sup> 1.4237, pK<sub>Est</sub> ~7.6**. The ester, prepared from *tert*-butyl azidoacetate by catalytic reduction (5% Pd/C and H<sub>2</sub>), has been purified *via* the phosphite salt. To the ester (23.6g, 0.18mole) in MeOH (150ml) is added phosphorous acid (15g, 0.18mole), the mixture is gently warmed to dissolve the latter and after cooling to 25°, Et<sub>2</sub>O (150ml) is added slowly and the stirred mixture is cooled at 0° for 12 hours. The *phosphite salt* is collected, filtered off, washed with Et<sub>2</sub>O and dried in an oven at 70° (32g, 82%, **m 144-147° dec.**). After recrystallisation from MeOH-isopropyl ether, the phosphite salt has **m 154-157° (dec)**. The phosphite salt (32g, 0.15mole) is added with stirring into aqueous 6N-sodium hydroxide solution (50ml) until all the solid has dissolved. The mixture is extracted with Et<sub>2</sub>O (2 x 20ml), the extract is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered, evaporated and the residue is distilled under vacuum to give the *tert*-butyl ester as an oil (14g, 72%). The *tert*-butyl group is versatile and is labile under acidic conditions which do not affect the blocked amino grouping. *Glycine tert-butyl ester hydrochloride* [27532-96-3] **M 167.6, has m 143° (EtOH/Et<sub>2</sub>O)**. [Vollmar & Dunn *J Org Chem* 25 387 1960, More and Rydon *Org Synth Coll Vol V* 586 1973, *Beilstein* 6 IV 2489.]

**Glycine hydrochloride** [6000-43-7] **M 111.5, m 176-178°, 185°, 187°**. Crystallise the salt from absolute EtOH or 80% EtOH. *Monoglycine hydrochloride* has **m 176-177°**, and *diglycine monohydrochloride* has **m 187°**. [Frost *J Am Chem Soc* 64 1286 1942, *Beilstein* 4 III 1111, 4 IV 2353.]

**Glycine methyl ester hydrochloride** [5680-79-5] **M 125.6, m 174°(dec), 177°(corrected), pK<sup>25</sup> 7.66**. Crystallise the ester salt from MeOH. [Werbin & Spoerri *J Am Chem Soc* 69 1682 1947, *Beilstein* 4 H 340, 4 III 1116.]

**Glycine *p*-nitrophenyl ester hydrobromide**. [7413-60-7] **M 277.1, m 214° (dec)**. Recrystallise the ester salt from MeOH by adding diethyl ether. [Alners et al. *Biochemical Preparations* 13 22 1971].

**Glycocyanine (*N*-guanylglycine)** [352-97-6] **M 117.1, m 280-284°(dec), >300°, pK<sup>25</sup> 2.86 (NH<sub>3</sub><sup>+</sup>)**. Recrystallise it from 15 parts of hot H<sub>2</sub>O, or by dissolving it in slightly more than the calculated amount of 2N HCl and precipitating it by adding an equivalent of 2N NaOH, filtering, washing with cold H<sub>2</sub>O and drying first *in vacuo*, then at 60° *in vacuo*. The *hydrochloride* has **m 200°(dec)** after recrystallisation from aqueous HCl as plates. The *picrate* forms needles from hot H<sub>2</sub>O with **m 210°(dec)**. [Brand & Brand *Org Synth Coll Vol III* 440 1955, Failey & Brand *J Biol Chem* 102 768 1933, King *J Chem Soc* 2375 1930, *Beilstein* 4 H 359, 4 I 477, 4 II 793, 4 III 1165.]

***N*-Glycylanilide** [555-48-6] **M 150.2, m 62°, pK<sub>Est</sub> ~8.0**. *N*-Glycylanilide crystallises from water as needles (*dihydrate*) and is soluble in Et<sub>2</sub>O. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp1915-1970 1961, *Beilstein* 4 H 343.]

**Glycylglycine** [556-50-3] **M 132.1, m 260-262°(dec), pK<sup>20</sup> 8.40, pK<sup>30</sup> 8.04, pK<sup>15</sup> 3.19, pK<sub>2</sub><sup>25</sup> 8.40**. Crystallise glycylglycine from aqueous 50% EtOH or water at 50-60° by addition of EtOH. Dry it at 110°. It sublimes at 190-200°/0.3mm with 30% recovery [Gross & Gradsky *J Am Chem Soc* 77 1678 1955, King *J Am Chem Soc* 79 6153 1957]. [*Beilstein* 4 IV 2459.]

**Glycylglycine hydrochloride** [13059-60-4] **M 168.6, m 215-220°, 235-236°, 260-262°, pK<sub>1</sub><sup>25</sup> 3.12, pK<sub>2</sub><sup>25</sup> 8.17**. Crystallise the salt twice from 95% EtOH. Single crystals are formed by slow evaporation of an aqueous

solution. [Parthasarathy *Acta Cryst (B)* **25** 509 1969, Mellon & Hoover *J Am Chem Soc* **73** 3879 1951, Garfinkel & Edsall *J Am Chem Soc* **80** 3818 1958, Beilstein **4** IV 2469.]

**Glycyl-L-proline** [704-15-4] M 172.2, m 184°(dec), 185°, 204°, pK<sub>1</sub><sup>25</sup> 2.81, pK<sub>2</sub><sup>25</sup> 8.65. Crystallise glycyl-L-proline from water at 50–60° by addition of EtOH. [Saidel *J Am Chem Soc* **77** 3893 1955, Bergmann et al. *Z Physiol Chem* **212** 79 1932, Beilstein **22** IV 49.]

**dl-Glycylserine** [687-38-7] M 162.2, m 197–199°(dec, sealed tube), 207°(dec), pK<sub>1</sub><sup>25</sup> 2.92, pK<sub>2</sub><sup>25</sup> 8.10. Crystallise it from H<sub>2</sub>O (charcoal) by addition of EtOH. [Fölsch *Acta Chem Scand* **12** 561–566 1958, NSR: Bovey & Tiers *J Am Chem Soc* **81** 2876 1959, Beilstein **4** III 1572, **4** IV 3140.]

**Gramicidin A (a pentadecapeptide from *Bacillus brevis*)** [11029-61-1] m ~229–230°(dec). Purify gramicidin A by countercurrent distribution from \*C<sub>6</sub>H<sub>6</sub>/CHCl<sub>3</sub>, MeOH/H<sub>2</sub>O (15:15:23:7) with 5000 tubes. Fractions are examined by UV (280nm) of small aliquots. Separation from gramicidin C and other material occurred after 999 transfers. [Gross & Witkop *Biochemistry* **4** 2495 1965, Bauer et al. *Biochemistry* **11** 3266 1972.] Purify it finally by recrystallisation from EtOH/H<sub>2</sub>O and dry it at 100°/10<sup>-2</sup>mm over KOH. It forms platelets m 229–230°. It is almost insoluble in H<sub>2</sub>O (0.6%) but soluble in lower alcohols, dry Me<sub>2</sub>CO, dioxane, acetic acid and pyridine. The commercial material is more difficult to crystallise than the synthetic compound. [Sarges & Witkop *J Am Chem Soc* **86** 1861 1964, **87** 2011, 2020 1965.] It has characteristic [α]<sub>D</sub><sup>20</sup> +27.3° (c 1.3, MeOH) and UV with λ<sub>max</sub> at 282nm (ε 22,100). The *N*-carbamoyldeformyl gramicidine A precipitates from EtOAc/petroleum ether (b 40–60°). [Beilstein **26** III/IV 4273.]

**Gramicidin C (a pentadecapeptide from *Bacillus brevis*)** [9062-61-7]. Purify as for Gramicidin A since they are isolated together and separated. [Sarges & Witkop *Biochemistry* **4** 2491 1965, Hunter & Schwartz “Gramicidins” in *Antibiotics I* (Gotlieb and Shaw Eds) Springer-Verlag, NY, p. 642 1967, as well as references above for Gramicidin A.]

**Gramicidin S** [113-73-5] M 1141.4, m 268–270°, [α]<sub>D</sub><sup>25</sup> -290° (c 0.5, EtOH + 30mM aqueous HCl {7:3}). Gramicidin S crystallises from EtOH. The *di-HCl* [15207-30-4] crystallises from EtOH (+ few drops of HCl) with m 277–278° (see below). [NMR: Gibbons et al. *Nature* **227** 840 1970, Beilstein **26** III/IV 4273.]

**Gramicidin S 2HCl (from *Bacillus brevis* Nagano)** [15207-30-4] M 1214.4, m 277–278°(dec), [α]<sub>D</sub><sup>24</sup> -289° (c 0.4, 70% H<sub>2</sub>O+EtOH). It crystallises in prisms from EtOH+aqueous HCl.

***N*-Guanyltiramine hydrochloride** [60-20-8] M 215.7, m 218°, pK<sub>1</sub> 10.2 (phenolic OH), pK<sub>2</sub> 12.4 (guanidino N). Purify the salt on a phosphocellulose column and elute with a gradient of aqueous NH<sub>3</sub> (0–10%). The second major peak has the characteristic tryptamine spectrum and is collected, and lyophilised to give white crystals of the *dihydrate* which dehydrate at 100°. It has UV with λ<sub>max</sub> at 274.5nm (ε 1,310) in 0.1N NaOH and 274.5nm (ε 1,330) at pH 7.0. Excitation λ<sub>max</sub> is at 280nm and emission λ<sub>max</sub> is at 330nm. [Mekalanos et al. *J Biol Chem* **254** 5849 1979.]

***S*-Histidine** [71-00-1] M 155.2, m 287°(dec), [α]<sub>D</sub><sup>25</sup> -39.7° (c 1, H<sub>2</sub>O), +13.0° (6M HCl), pK<sub>1</sub><sup>25</sup> 1.96, pK<sub>2</sub><sup>25</sup> 6.12, pK<sub>3</sub><sup>25</sup> 9.17. A likely impurity is arginine. *S*-Histidine is adsorbed from aqueous solution onto a Dowex 50-H<sup>+</sup> ion-exchange resin, washed with 1.5M HCl (to remove other amino acids), then eluted with 4M HCl as the *dihydrochloride*. This purified *dihydrochloride* (see below) is finally dissolved in water, the pH adjusted to 7.0, and the *free zwitterionic base* crystallises out on addition of EtOH. Its solubility in H<sub>2</sub>O is 4.2% at 25°. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 1971–1993 1961, Beilstein **25** III/IV 4344.]

***S*-Histidine dihydrochloride** [1007-42-7] M 242.1, m 245°, [α]<sub>D</sub><sup>20</sup> +47.5° (c 2, H<sub>2</sub>O). The dihydrochloride crystallises from water or aqueous EtOH and is washed with acetone, then diethyl ether. *Alternatively*, convert it to the histidine *di-(3,4-dichlorobenzenesulfonate)* salt by dissolving 3,4-dichlorobenzenesulfonic acid (1.5g/10ml) in the aqueous histidine solution with warming, and then the solution is cooled in ice. The resulting crystals (m 280° dec) can be recrystallised from 5% aqueous 3,4-dichlorobenzenesulfonic acid, then dried over

CaCl<sub>2</sub> under vacuum, and washed with diethyl ether to remove excess reagent. The dihydrochloride can be regenerated by passing the solution through a Dowex-1 (Cl<sup>-</sup> form) ion-exchange column. The solid is obtained by evaporation of the solution on a steam bath or better in a vacuum. [Greenstein & Winitz, *The Amino Acids* Vol 3 p 1976 1961, *Beilstein* 25/16 V 366.]

**S-Histidine monohydrochloride (H<sub>2</sub>O)** [5934-29-2 (H<sub>2</sub>O), 7048-02-4] M 209.6, m 80° monohydrate, 254°(dec, anhydrous), [α]<sub>D</sub><sup>25</sup> +13.0° (6M HCl). Crystallise the mono-hydrochloride from aqueous EtOH or 60% aqueous EtOH (m 259°dec). Alternatively, dissolve 10g in 50ml of H<sub>2</sub>O, decolourise with Norite, filter, evaporate it in a vacuum to a syrup, cool to room temperature, add 95% EtOH with stirring until slightly turbid, scratch the sides of the vessel until crystals form, then add slowly 40ml of EtOH and keep at 0° overnight, filter the solid off, wash it several times with EtOH and dry it in a vacuum. [Cox *J Biol Chem* 78 475 1929, Cox et al. *J Biol Chem* 81 73 1929, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 1972, 2098 1961, *Beilstein* 25 II 407, 25 III/IV 4346.]

**L-Homocysteine (2-amino-4-mercaptoputyric acid)** [6027-13-0] M 135.2, m 232-233°, [α]<sub>D</sub><sup>25</sup> +153° (c 13, 5N HCl), pK<sub>1</sub><sup>25</sup> 2.22 (CO<sub>2</sub>H), pK<sub>2</sub><sup>25</sup> 8.87 (NH<sub>3</sub><sup>+</sup>), pK<sub>3</sub><sup>25</sup> 10.86 (SH). Crystallise L-homocysteine from aqueous EtOH. All operations should be carried out under N<sub>2</sub> as the thiol readily oxidizes in air. The acid (3g) is dissolved in freshly boiled H<sub>2</sub>O (30ml) under N<sub>2</sub>, cooled under N<sub>2</sub> (all operations should be under N<sub>2</sub>), add absolute EtOH (100ml), the acid is filtered off, and a second crop is obtained by diluting the filtrate to 500ml with absolute EtOH, kept overnight in a refrigerator, filtering, washing with EtOH and drying in a vacuum. Store it under N<sub>2</sub> or argon.

The S-benzyl derivative is repeatedly crystallised from H<sub>2</sub>O, or by dissolving it in HCl followed by slow addition of ammonia. It has m 240-241°, [α]<sub>D</sub><sup>25</sup> +27° (c 13, 5N HCl). [Riedel & du Vigneaud *J Biol Chem* 112 149 1935, du Vigneaud & Patterson *J Biol Chem* 109 101 1935, du Vigneaud *J Biol Chem* 126 217 1938, du Vigneaud & Brown *Biochemical Preparations* 5 93, 95 1975, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2667-2670 1961, *Beilstein* 4 IV 3189, IR: Koegel et al. *J Am Chem Soc* 77 5708 1955.]

**dl-Homocysteine (2-amino-4-mercaptoputyric acid)** [454-29-5] M 135.2, m 234-235°(corr, dec). Purify it as for the L-isomer. [Allen & Steinmann *J Am Chem Soc* 74 3932 1952, and references for the L-isomer above, *Beilstein* 4 IV 3189.]

**dl-Homocystine** [462-10-2, 870-93-9 (±)] M 268.4, m 263-265°(dec), pK<sub>1</sub><sup>25</sup> 1.59 (CO<sub>2</sub>H), pK<sub>2</sub><sup>25</sup> 2.54 (CO<sub>2</sub>H), pK<sub>3</sub><sup>25</sup> 8.52 (NH<sub>3</sub><sup>+</sup>), pK<sub>4</sub><sup>25</sup> 9.44 (NH<sub>3</sub><sup>+</sup>). dl-Homocystine crystallises in platelets from water with 1H<sub>2</sub>O and m 258-260°(dec), all operations should be carried out under N<sub>2</sub>. [Sudo *J Chem Soc Jpn (Pure Chem Sect)* 79 81, 86, 87 1958, *Beilstein* 4 IV 3199.]

**L(S,S)-Homocystine** [626-72-2] M 268.4, m 281-284°(dec), [α]<sub>D</sub><sup>26</sup> +79° (c 1, M HCl), [α]<sub>D</sub><sup>21</sup> -16° (c 0.06, H<sub>2</sub>O), pK (see above). The acid (3g) is dissolved in freshly boiled H<sub>2</sub>O (30ml) under N<sub>2</sub>, cooled under N<sub>2</sub> (all operations should be under N<sub>2</sub>), absolute EtOH (100ml) is added, the acid is filtered off, and a second crop is obtained by diluting the filtrate to 500ml with absolute EtOH, kept overnight in a refrigerator, filtered, washed with EtOH and dried in a vacuum. The D(R,R)-form has similar properties but is -ve in M HCl and +ve in H<sub>2</sub>O. [du Vigneaud & Patterson *J Biol Chem* 109 101 1935, du Vigneaud & Brown *Biochemical Preparations* 5 93, 95 1975, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2667-2670 1961, *Beilstein* 4 III 1643, 4 IV 3199, Koegel et al. *J Am Chem Soc* 77 5708 1955.]

**L-Homoserine (2-amino-4-hydroxybutyric acid)** [672-15-1] M 119.1, m 203°, [α]<sub>D</sub><sup>26</sup> +18.3° (in 2M HCl), pK<sub>Est(1)</sub> ~2.1, pK<sub>Est(2)</sub> ~9.3. Likely impurities are N-chloroacetyl-L-homoserine, N-chloroacetyl-D-homoserine, L-homoserine, homoserine lactone, homoserine anhydride (formed in strong solutions of homoserine if slightly acidic). It cyclises to the lactone in strongly acidic solution. It crystallises from water by adding 9 volumes of EtOH. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2612-2616 1961, *Beilstein* 4 IV 3187.]

**erythro-3-Hydroxy-*RS*-aspartic acid** [6532-76-9] **M 149.1, pK<sub>1</sub><sup>25</sup> 1.91, pK<sub>2</sub><sup>25</sup> 3.51, pK<sub>3</sub><sup>25</sup> 9.11.** Likely impurities are 3-chloromalic acid, ammonium chloride, *threo*-3-hydroxyaspartic acid. Crystallise it from water. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 1** p 214, **Vol 3** p 2416 1961.]

**β-Hydroxyglutamic acid** [533-62-0] **M 163.1, m 100°(dec), pK<sub>1</sub><sup>25</sup> 2.27, pK<sub>2</sub><sup>25</sup> 4.29, pK<sub>3</sub><sup>25</sup> 9.66.** Crystallise the acid from water. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 1** pp 211-213, **Vol 3** p 2422 1961.]

**5-Hydroxy-L-lysine monohydrochloride** [32685-69-1; 13204-98-3] **M 198.7, m 225°(dec), [α]<sub>D</sub><sup>25</sup> +17.8° (6M HCl), pK<sub>2</sub><sup>25</sup> 8.85, pK<sub>3</sub><sup>25</sup> 9.83.** Likely impurities are 5-*allo*-hydroxy-(D and L)-lysine, histidine, lysine, ornithine. Crystallise the hydrochloride from water by adding 2-9 volumes of EtOH stepwise. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2000-2009 1961.]

**DL-erythro-3-Hydroxynorvaline (2-amino-3-hydroxypentanoic acid)** [34042-00-7] **M 133.2, m 257-259°(dec), 263°(dec), pK<sub>1</sub><sup>20</sup> 2.32, pK<sub>2</sub><sup>20</sup> 9.12.** Purify it by recrystallisation from aqueous EtOH. The *Cu* salt has **m 255-256° (dec)**, the *benzoyl* derivative has **m 181°**, and the *N*-phenylcarbamoyl derivative has **m 164°**. [Buston et al. *J Biol Chem* **204** 665 1953, *Beilstein* **4** IV 3220.]

***N*-(*p*-Hydroxyphenyl)glycine** [22818-40-2] **M 167.2, m >240°(dec), [α]<sub>D</sub><sup>20</sup> -156° (c 1, M HCl), pK<sub>Est(1)</sub>~2, pK<sub>Est(2)</sub>~4.5, pK<sub>Est(3)</sub>~10.3.** Crystallise it from water and dry it *in vacuo*. [*Beilstein* **14** I 659.]

***trans*-L-4-Hydroxyproline** [51-35-4] **M 131.1, m 274°, [α]<sub>D</sub><sup>20</sup> -76.0° (c 5, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 1.86, pK<sub>2</sub><sup>25</sup> 9.79.** Crystallise it from MeOH/EtOH (1:1). Separation from normal *allo*-isomer can be achieved by crystallisation of the copper salts [see Levine *Biochemical Preparations* **8** 114 1961]. Separation from proline is achieved via the crystalline picrate, CdCl<sub>2</sub>, or acid ammonium rhodanate salts [see Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2182 1961, Kapfhammer & Mohn *Z Physiol Chem* **306** 76 1956]. [*Beilstein* **22/5** V 7.]

**5-Hydroxy-L-tryptophan** [4350-09-8] **M 220.2, m 273°(dec), [α]<sub>D</sub><sup>22</sup> -32.5°, [α]<sub>D</sub><sup>20</sup> -73.5° (c 1, H<sub>2</sub>O), pK<sub>Est(1)</sub>~2.4, pK<sub>Est(2)</sub>~9.0, pK<sub>Est(3)</sub>~9.4, pK<sub>Est(4)</sub> 16 (NH).** Likely impurities are 5-hydroxy-D-tryptophan and 5-benzyloxytryptophan. Crystallise 5-hydroxy-L-tryptophan under nitrogen from water by adding EtOH. Store it under nitrogen. Also dissolve it in the minimum volume of hot H<sub>2</sub>O (~0.7g in 4ml) under nitrogen (charcoal) and allowed it to crystallise at 5°. The *picrolonate* crystallises from H<sub>2</sub>O with **m 184-186°(dec)**. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2732-2737 1961, Morris & Armstrong *J Org Chem* **22** 306 1957, *Beilstein* **22/14** V 278.]

**(±)-Ibotenic acid monohydrate (α-[3-hydroxy-5-isoxazolyl]-glycine, α-amino-3-hydroxy-5-isoxazoleacetic acid)** [2552-55-8] **M 176.1, m 144-146° (monohydrate), 151-152° (anhydrous), 148-151°, pK<sub>1</sub> 2, pK<sub>2</sub> 5.1, pK<sub>3</sub> 8.2.** It has been converted to the ammonium salt (**m 121-123° dec**) dissolved in H<sub>2</sub>O, passed through an Amberlite IR 120 resin (H<sup>+</sup> form) and eluted with H<sub>2</sub>O. The acidic fractions are collected, evaporated to dryness and the residue recrystallises from H<sub>2</sub>O as the *monohydrate* (**m 144-146°**). The anhydrous acid is obtained by making a slurry with MeOH, decanting and evaporating to dryness, and repeating the process twice more to give the *anhydrous acid* (**m 151-152°**). Recrystallisation from H<sub>2</sub>O gives the *monohydrate*. [Nakamura *Chem Pharm Bull Jpn* **19** 46 1971.] The *ethyl ester* forms needles when crystallised from a small volume of Et<sub>2</sub>O and has **m 78-79°** and IR (CHCl<sub>3</sub>) with  $\nu_{\max}$  □□ 3500-2300 (OH), 1742 (ester C=O), 1628, 1528cm<sup>-1</sup>, and UV with  $\lambda_{\max}$  □□ (EtOH) at 206nm (ε 7,080). The *hydrazide* has **m 174-175°** (from MeOH) with IR (KBr) 1656 (C=O)cm<sup>-1</sup>.

**Iminodiacetic acid** [142-73-4] **M 133.1, m 225°(dec), 240°(dec), 247.5°(dec), b 126-127°/14mm, pK<sub>1</sub><sup>25</sup> 2.50, pK<sub>2</sub><sup>25</sup> 9.40.** Crystallise the acid several times from water. The *N*-Methyl derivative **m 215°** is purified by dissolving it in an equal weight of warm H<sub>2</sub>O and adding 3 volumes of MeOH [Kiematsu et al. *Org Synth Coll Vol II* 397 1943]. [Laberek & Martell *J Am Chem Soc* **74** 5052 1952, *Beilstein* **4** III 2428, **4** IV 1176.]

**3-Iodo-L-tyrosine** [70-78-0] M 307.1, m 205-208°(dec),  $[\alpha]_{\text{D}}^{25} -4.4^{\circ}$  (c 5, 1M HCl),  $\text{pK}_{\text{Est}(2)} \sim 2.1$ ,  $\text{pK}_{\text{Est}(3)} \sim 6.4$ ,  $\text{pK}_{\text{I}}^{25} 8.7$ . Likely impurities are tyrosine, diiodotyrosine and iodide. Crystallise it by dissolving it in concentrated ammonia (~200mg in ~20ml), evaporate to ~5ml and  $\text{NH}_4\text{Cl}$  is added to pH4.5—5.0. After a few hours at 0°, the amino acid crystallises in needles. It is filtered off, washed with a little ice-cold  $\text{H}_2\text{O}$  and dried in a vacuum. *Alternatively*, dissolve it in dilute ammonia at room temperature, then add dilute acetic acid to pH 6. Store it at 0°. Recrystallisation of ~250mg from  $\text{H}_2\text{O}$  (~5ml) removes any diiodotyrosine. It is an inhibitor of tyrosine hydroxylase with a  $K_i$  of ~500nM. [Harrington & Rivers *Biochem J* 38 320 1944, Rivers *Chem & Ind (London)* 21 1956, *Beilstein* 14 III 1562, 14 IV 1562.]

**L-Isoleucine** [73-32-5] M 131.2, m 285-286°(dec),  $[\alpha]_{\text{D}}^{20} +40.6^{\circ}$  (6M HCl)  $\text{pK}_{\text{I}}^{25} 2.66$ ,  $\text{pK}_{\text{I}}^{25} 9.69$ . Crystallise L-isoleucine from  $\text{H}_2\text{O}$  by addition of 4 volumes of EtOH or from aqueous MeOH. It sublimes at 170-181°/0.3mm with 99.7% recovery, unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 1 p 183-191, Vol 3 pp 2043-2073 1961, Huffman & Ingersoll *J Am Chem Soc* 73 3366 1951, *Beilstein* 4 IV 2775.]

**DL-Isoserine** ( $\pm$ -3-amino-2-hydroxypropionic acid) [632-12-2] M 105.1, m 235°(dec), 237°(dec), 245°(dec), 250-252°(dec),  $\text{pK}_{\text{I}}^{25} 2.78$  (acidic),  $\text{pK}_{\text{I}}^{25} 9.27$  (basic). Recrystallise it from  $\text{H}_2\text{O}$  or 50% aqueous EtOH. It has an isoelectric pH of 6.02. [Rinderknecht & Niemann *J Am Chem Soc* 75 6322 1953, Gundermann & Holtmann *Chem Ber* 91 160 1958, Emerson et al. *J Biol Chem* 92 451 1931.] The *hydrobromide* has m 128-130° (from aqueous HBr) [Schöberl & Braun *Justus Liebigs Ann Chem* 542 288 1939]. [*Beilstein* 4 H 503, 4 IV 3116.]

**L-Isovaline** (2-amino-2-methylbutyric acid) [595-40-4] M 117.2, m ca 300° (sublimes in vac),  $[\alpha]_{\text{D}}^{25} +113.1^{\circ}$  (c 5,  $\text{H}_2\text{O}$ ),  $[\alpha]_{\text{D}}^{25} +10^{\circ}$  (5M HCl),  $\text{pK}_{\text{Est}(1)} \sim 2.4$ ,  $\text{pK}_{\text{Est}(2)} \sim 9.7$ . Crystallise it from aqueous  $\text{Me}_2\text{CO}$ , or better by dissolving in  $\text{H}_2\text{O}$  and adding excess  $\text{Me}_2\text{CO}$ . [Baker et al. *J Am Chem Soc* 74 4701 1952, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 2573-2577 1961.]

**L-Leucine** [61-90-5] M 131.2, m 293-295°(dec),  $[\alpha]_{\text{D}}^{25} +15.6^{\circ}$  (5M HCl),  $\text{pK}_{\text{I}}^{25} 2.33$ ,  $\text{pK}_{\text{I}}^{25} 9.74$ . Likely impurities are isoleucine, valine, and methionine. Crystallise L-leucine from water by adding 4 volumes of EtOH. It sublimes at 180-188°/0.3mm with 99.1% recovery, and unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 p 2075-2094 1961, Kameda et al. *J Pharm Soc Jpn* 78 763 1958, *Beilstein* 4 IV 2738.]

**L-Lysine** [56-87-1] M 146.2, [39665-12-8 monohydrate] M 164.2, m >210°(dec),  $[\alpha]_{\text{D}}^{20} +25^{\circ}$  (c 2, 6M HCl),  $\text{pK}_{\text{I}} 2.18$ ,  $\text{pK}_{\text{I}} 8.95$ ,  $\text{pK}_{\text{I}} 10.53$ . Crystallise L-lysine from aqueous EtOH. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2097-2122 1961, Kearley & Ingersoll *J Am Chem Soc* 73 5783 1951, *Beilstein* 4 IV 2717.]

**L-Lysine dihydrochloride** [657-26-1] M 219.1, m 193°, 199-201°, 203-204°,  $[\alpha]_{\text{D}}^{25} +25.9^{\circ}$  (5M HCl). Crystallise it from MeOH, in the presence of excess HCl, by adding diethyl ether. [Yoneya *J Biochem (Tokyo)* 38 343 1951, Kearley & Ingersoll *J Am Chem Soc* 73 5783 1951, *Beilstein* 4 IV 2717.]

**L-Lysine monohydrochloride** [657-27-2] M 182.7, m 256°(dec),  $[\alpha]_{\text{D}}^{25} +20.5^{\circ}$  (c 5, 5M HCl). Likely impurities are arginine, D-lysine, 2,6-diaminoheptanedioic acid and glutamic acid. Crystallise the monohydrochloride from water at pH 4-6 by adding 4 volumes of EtOH. At above 60% relative humidity it forms a *dihydrate*. [Birhbaum et al. *J Biol Chem* 194 455, 468 1952, Kearley & Ingersoll *J Am Chem Soc* 73 5783 1951, *Beilstein* 4 IV 2717.]

**$\alpha$ -Melanotropin** [581-05-5] (a tridecapeptide,  $\alpha$ -MSH, melanocyte stimulating hormone), M 1664.9,  $[\alpha]_{\text{D}}^{25} -58.5^{\circ}$  (c 0.4, 10% aqueous AcOH). Its solubility in  $\text{H}_2\text{O}$  is 1mg/ml. It is separated from the extract by ion-exchange on carboxymethyl cellulose, desalted, evaporated and lyophilised, then chromatographed on

Sephadex G-25. [Lande et al. *Biochemical Preparations* **13** 45 1971.]

**$\beta$ -Melanotropin** [9034-42-8] (octadeca to docosa peptides), amorphous. An extract of  $\beta$ -melanotropin is purified by ion-exchange on carboxymethyl cellulose, desalted, evaporated and lyophilised, then chromatographed on Sephadex G-25. [Lande et al. *Biochemical Preparations* **13** 45 1971.]

**Melphalan** (4-[bis-{2-chloroethyl}amino]-L-phenylalanine) [148-82-3] **M 305.2, m 182-183° (dec), 183-185°,  $[\alpha]_D^{25} +7.5^\circ$  (c 1.33, 1.0 N HCl),  $[\alpha]_D^{20} -28^\circ$  (c 0.8, MeOH),  $pK_{Est} \sim 6.4$** . Purify melphalan by recrystallisation from MeOH, and its solubility is 5% in 95% EtOH containing one drop of 6N HCl. It is soluble in EtOH and propylene glycol but is almost insoluble in H<sub>2</sub>O. The *RS*-form has **m 180-181°**, and the *R*-form crystallises from MeOH with **m 181.5-182°** and  **$[\alpha]_D^{21} -7.5^\circ$  (c 1.26, 1.0 N HCl)**. [Bergel & Stock *J Chem Soc* 2409 1954, *Beilstein* **14** IV 1689.]

***dl*-Methionine** (*RS*-2-amino-4-methylthiobutyric acid) [59-51-8] **M 149.2, m 281°(dec),  $pK_1^{25} 2.28$ ,  $pK_2^{25} 9.21$** . Crystallise it from hot water or EtOH. Also purify it by dissolving it in H<sub>2</sub>O and passing through an Amberlite IR-120 resin (NH<sub>4</sub><sup>+</sup> form). The eluate is concentrated and then passed through Amberlite IR-4B resin, and this eluate is evaporated to dryness. The residue is washed with EtOH, then Me<sub>2</sub>CO, dried and recrystallised from aqueous EtOH (colourless plates) [Baddiley & Jamieson *J Chem Soc* 4283 1954]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2125 1961, *Beilstein* **4** IV 3190.]

**L-Methionine** [63-68-3] **M 149.2, m 277-279°(dec), 283°(dec),  $[\alpha]_D^{25} +21.2^\circ$  (0.2M HCl),  $pK_1^{25} 2.13$ ,  $pK_2^{25} 9.73$** . Crystallise L-methionine from aqueous EtOH. Also purify it by dissolving ~0.5g of amino acid in ~10ml of hot H<sub>2</sub>O, filtering, adjusting the pH to 5.8 with 5N HCl, collecting the solid after addition of ~20ml of EtOH. It is recrystallised by dissolving in H<sub>2</sub>O and adding EtOH. It sublimes at 197-208°/0.3mm with 99.8% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. [Milne & Peng *J Am Chem Soc* **79** 647 1957, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2125-2152 1961, *Beilstein* **4** IV 3189.]

***dl*-Methionine sulfoxide** [454-41-1, 62697-73-8] **M 165.2, m >240°(dec), 241-242°(dec)**. Likely impurities are *dl*-methionine sulfone and *dl*-methionine. Crystallise the sulfoxide by dissolving it in hot H<sub>2</sub>O and adding excess EtOH. [Lepp & Dunn *Biochemical Preparations* **4** 80 1955, Micheel & Schmitz *Chem Ber* **72** 518 1939, *Beilstein* **4** III 1650, **4** IV 3192.]

**S-Methyl-L-cysteine** [1187-84-4] **M 135.2, m 207-211°, ~240°(dec), 267-270°,  $[\alpha]_D^{20} -32.0^\circ$  (-34.5°) (c 5, H<sub>2</sub>O),  $pK_1^{25} 1.94$  (COOH),  $pK_2^{25} 8.73$  (NH<sub>2</sub>, 8.97)**. Likely impurities are cysteine and S-methyl-*dl*-cysteine. Crystallise it from H<sub>2</sub>O by adding 4 volumes of EtOH. It also crystallises from MeOH with **m 234-236°(dec)**, but after sublimation it has **m 267-270°** and  **$[\alpha]_D^{27} -31.6^\circ$  (c 1, H<sub>2</sub>O)**. [Rinderknecht et al. *Helv Chim Acta* **41** 1, 10 1958, Theodoropoulos *Acta Chem Scand* **13** 383 1959, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 1904 1961, *Beilstein* **4** IV 3145.]

**$\alpha$ -Methylmethionine** [562-48-1] **M 163.0, m 283-284°,  $pK_{Est(1)} \sim 2.1$ ,  $pK^{30} 9.45$** . Crystallise  $\alpha$ -methylmethionine from aqueous EtOH or H<sub>2</sub>O. [Pfister et al. *J Am Chem Soc* **77** 697 1955, Potts *J Chem Soc* 1623 1955, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2566 1961.]

**S-Methyl-L-methionine chloride** See Vitamin U in "Miscellaneous Compounds", this chapter.

***N*-Methyltryptophan** (L-abrine) [526-31-8] **M 218.3, m 295°(dec with darkening and sintering),  $[\alpha]_D^{21} +44.4^\circ$  (c 2.8, 0.5M HCl),  $[\alpha]_D^{20} +65^\circ$  (c 1, 0.5N NaOH),  $pI 10.10$ ,  $pK_{Est(1)} \sim 2.3$ ,  $pK_{Est(2)} \sim 9.7$** . Crystallise L-abrine from H<sub>2</sub>O or EtOH/H<sub>2</sub>O mixture and dry it for 2 days at 60° in high vacuum; it has **m 275-290°(dec with browning at 230°)** and  **$[\alpha]_D^{21} +47.2^\circ$  (c 2, 0.5N HCl)** [Peter et al. *Helv Chim Acta* **46** 577 1963]. [Gregory & Morley *J Chem Soc* 913 1968, *Beilstein* **22/14** V 40.]

***dl*-5-Methyltryptophan** [951-55-3] **M 218.3, m 275°(dec) [ $pK$  see tryptophan]**. Crystallise *dl*-5-methyltryptophan from aqueous EtOH after dissolving it in aqueous NaOH, precipitating with AcOH, filtering

the solid off and drying for 24 hours at 50°. [Jackman & Archer *J Am Chem Soc* **68** 2105 1946, *Beilstein* **22** IV 6815.] The *picrate* crystallises from MeOH with **m** 202°(dec). The *N*-phenylcarbamoyl derivative crystallises from aqueous MeOH with **m** 202°. [Gordon & Jackson *J Biol Chem* **110** 151, 154 1935.]

**Nisin** [1414-45-5] **M 3354.2**. This polypeptide from *S. lactis* is purified by crystallisation from 80% (v/v) EtOH and by countercurrent distribution. The synthetic polypeptide antibiotic can also be purified by preparative HPLC and assayed by HPLC on a Nucleosil 3007C<sub>18</sub> (6 x 250mm) column using a MeCN—0.01M HCl gradient (30-50%), at 2%/minute, and flow rate of 1.5ml/minute to give a retention time of 8.1 minutes; or MeCN—0.3M guanidine-HCl gradient (30-50%), at 2%/minute, and flow rate of 1.5ml/minute to give a retention time of 10.9 minutes. FAB-MS gave the pseudomolecular ion *m/z* at 3352.7 (M + H)<sup>+</sup>. It is soluble in dilute acid and is stable even on boiling. [Berridge et al. *Biochem J* **52** 529 1952, synthesis by Fukase et al. *Tetrahedron Lett* **29** 795 1988.]

**Norleucine (α-amino-*n*-caproic acid)** [*R*(+) 327-56-0, *S*(-) 327-57-1] **M 117.2, m 301°**, [α]<sub>D</sub><sup>20</sup> (+) and (-) 28° (c 5, 5M HCl), [*RS*: 616-06-8] **m 297-300°** (sublimes partially at ~280°), **pK<sub>1</sub> 2.39, pK<sub>2</sub> 9.76** (for *RS*). Crystallise norleucine from water or aqueous MeOH. [Huffman & Ingersoll *J Am Chem Soc* **73** 3366 1951, *Beilstein* **4** III 1386, **4** IV 2628.]

**Norvaline (R-α-amino-*n*-valeric acid)** [*R*(+) 2031-12-9, *S*(-) 6600-40-4] **M 117.2, m 305°(dec)**, [α]<sub>D</sub><sup>20</sup> (+) and (-) 25° (c 10, 5M HCl), **pI 6.04, pK<sub>1</sub><sup>25</sup> 2.36, pK<sub>2</sub><sup>25</sup> 9.87** (9.72). Crystallise norvaline from aqueous EtOH or water. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2390-2399 1961, *Beilstein* **4** III 1331-1333, **4** IV 128, 2629.]

**L-Ornithine** [70-26-8] **M 132.2, m 140°**, [α]<sub>D</sub><sup>25</sup> +16° (c 0.5, H<sub>2</sub>O), **pK<sub>1</sub><sup>20</sup> 2.11, pK<sub>2</sub><sup>20</sup> 8.39, pK<sub>3</sub><sup>25</sup> 10.59**. Crystallise L-ornithine from water containing 1mM EDTA (to remove metal ions). [Perrin *J Chem Soc* 3125 1958, Rivard *Biochemical Preparations* **3** 97 1955, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2477-2491 1961, *Beilstein* **4** III 1346, **4** IV 2644.]

**L-Ornithine monohydrochloride** [3184-13-2] **M 168.6, m 230-232°(dec), 233°(dec), 236.5-237.5°(dec)**, [α]<sub>D</sub><sup>25</sup> +28.3° (5M HCl). Likely impurities are citrulline, arginine and D-ornithine. Crystallise the monohydrochloride from water by adding 4 volumes of EtOH and dry it in a vacuum desiccator over fused CaCl<sub>2</sub>. [Rivard *Biochemical Preparations* **3** 98 1955.] The *dihydrochloride* [6211-16-1] has **m 202-203°** and [α]<sub>D</sub><sup>20</sup> +18.4° (c 2.3, 6N HCl) after recrystallisation from MeOH/Et<sub>2</sub>O [Zaoral & Rudinger *Col Czech Chem Commun* **24** 2009 1959]. [*Beilstein* **4** IV 2644.]

**Oxytocin** [50-56-6] **M 1007.2, m dec on heating**, [α]<sub>D</sub><sup>22</sup> -26.2° (c 0.53, N AcOH). It is a cyclic nonapeptide which is purified by countercurrent distribution between solvent and buffer. It is soluble in H<sub>2</sub>O, *n*-BuOH and isoBuOH. [Bodanszky & du Vigneaud *J Am Chem Soc* **81** 2504 1959, Cash et al. *J Med Pharm Chem* **5** 413 1962, Sakakibara et al. *Bull Chem Soc Jpn* **38** 120 1965; solid phase synthesis: Bayer & Hagenmyer *Tetrahedron Lett* 2037 1968.] It was also synthesised on a solid phase matrix and finally purified as follows: A Sephadex G-25 column is equilibrated with the aqueous phase of a mixture of 3.5% AcOH (containing 1.5% of pyridine)/*n*-BuOH/\*C<sub>6</sub>H<sub>6</sub> (2:1:1) and then the organic phase of this mixture is run through. A solution of oxytocin (100mg) in H<sub>2</sub>O (2ml) is applied to the column which is then eluted with the organic layer of the above mixture. The fractions containing the major peak [as determined by the Folin-Lowry protein assay: Fryer et al. *Anal Biochem* **153** 262 1986] are pooled, diluted with twice their volume of H<sub>2</sub>O, evaporated to a small volume and lyophilised to give oxytocin as a pure white powder (20mg, 508 U/mg). [Ives *Can J Chem* **46** 2318 1968, *Beilstein* **22** III/IV 82.]

***dl*-Phenylalanine** [150-30-1] **M 165.2, m 265-266°(capillary, dec), 271-273°(dec), 282-284°(dec)**, **pK<sub>1</sub><sup>25</sup> 2.58, pK<sub>2</sub><sup>25</sup> 9.24**. *dl*-Phenylalanine crystallises from H<sub>2</sub>O or H<sub>2</sub>O/EtOH in large plates and is dried under vacuum over P<sub>2</sub>O<sub>5</sub>. *S*-Phenylalanine ethyl ester hydrochloride [3182-93-2] has **m 156-158°** and [α]<sub>D</sub><sup>20</sup> -7.8° (c

2, H<sub>2</sub>O) after crystallisation from EtOH/Et<sub>2</sub>O [Billimoria & Cook *J Chem Soc* 2328 1949, *Beilstein* 14 IV 1556]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2156-2175 1961, *Beilstein* 14 III 1229, 14 IV 1553.]

**L-Phenylalanine** [63-91-2] M 165.2, m 280°(dec), 281-183°(dec),  $[\alpha]_D^{25}$  -34.0° (c 2, H<sub>2</sub>O). Likely impurities are leucine, valine, methionine and tyrosine. Crystallise L-phenylalanine from water by adding 4 volumes of EtOH. Dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub>. Also crystallise it from saturated refluxing aqueous solutions at neutral pH, or 1:1 (v/v) EtOH/water solution, or conc HCl. It sublimes at 176-184°/0.3mm with 98.7% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2156-2175 1961, *Beilstein* 14 IV 1552.]

**L- $\alpha$ -Phenylglycine** [2935-35-5] M 151.2, m 305-310°, 305-308°(capillary, dec),  $[\alpha]_{546}^{25}$  +188° (c 1, M HCl), pK<sub>1</sub><sup>25</sup> 1.83, pK<sub>2</sub><sup>25</sup> 4.39 (for *dl*). Crystallise it from EtOH. [Kaneko *J Chem Soc Jpn* 60 538 1939, Rudman et al. *J Am Chem Soc* 74 551 1952, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2694-2697 1961, *Beilstein* 14 III 1187, 14 IV 1317.]

**Phenylglycine-*o*-carboxylic acid [N-(2-carboxyphenyl)glycine]** [612-42-0, 64241-57-2 Me Ester, 67990-19-6 mono-Na Salt, 71807-57-3 di-Na Salt] M 195.2, m 206°, 208°, 220°, pK<sub>1</sub><sup>20</sup> 5.44 (CO<sub>2</sub>H), pK<sub>2</sub><sup>20</sup> 6.96 (CO<sub>2</sub>H) (in 50% aqueous dioxane). Crystallise the acid from hot water (charcoal). It forms complexes with Cu<sup>2+</sup>, Zn<sup>2+</sup>, Cd<sup>2+</sup>, Co<sup>2+</sup> and Ni<sup>2+</sup> in aqueous dioxane. [Roileanu et al. *Rev Roumaine Chim* 12 105 1967, *Aldrich library of* <sup>13</sup>C, <sup>1</sup>H FTNMR Spectra, NMR 2 1181A, *Beilstein* 14 H 348, 14 I 544, 14 II 225, 14 III 938.]

**D-Pipecolinic acid (R-piperidine-2-carboxylic acid)** [1723-00-8] M 129.2, m 264°(dec), 267°(dec), ~280°(dec),  $[\alpha]_D^{25}$  +26.2° (c 2, H<sub>2</sub>O),  $[\alpha]_D^{25}$  +35.7° (H<sub>2</sub>O), pK<sub>1</sub><sup>20</sup> 2.29 (CO<sub>2</sub>H), pK<sub>2</sub><sup>20</sup> 10.77 (NH<sup>+</sup>). D-Pipecolinic acid recrystallises as platelets from EtOH and is soluble in H<sub>2</sub>O. The *hydrochloride* has m 256-257°(dec) from H<sub>2</sub>O and  $[\alpha]_D^{25}$  +10.8° (c 2, H<sub>2</sub>O). [*cf* p 422, Lukés et al. *Col Czech Chem Commun* 22 286 1957, Bayerman *Rec Trav Chim Pays Bas* 78 134 1959, Asher et al. *Tetrahedron Lett* 22 141 1981, *Beilstein* 22/1 V 220.]

**L-Pipecolinic acid (S-piperidine-2-carboxylic acid)** [3105-95-1] M 129.2, m 268°(dec), 271°(dec), ~280°(dec),  $[\alpha]_D^{20}$  -26° (c 4, H<sub>2</sub>O),  $[\alpha]_D^{25}$  -34.9° (H<sub>2</sub>O). Recrystallise L-pipecolinic acid from aqueous EtOH, and it sublimes as needles in a vacuum. It is sparingly soluble in absolute EtOH, Me<sub>2</sub>CO or CHCl<sub>3</sub> but insoluble in Et<sub>2</sub>O. The *hydrochloride* has m 258-259°(dec, from MeOH) and  $[\alpha]_D^{25}$  -10.8° (c 10, H<sub>2</sub>O). [Fuji & Myoshi *Bull Chem Soc Jpn* 48 1241 1975, *Beilstein* 22/1 V 220.]

**Piperidine-4-carboxylic acid (isonipecotic acid)** [498-94-2] M 129.2, m 336°(dec, darkens at ~300°), pK<sub>Est(1)</sub> ~ 4.3 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~ 10.6 (NH<sup>+</sup>). It crystallises from H<sub>2</sub>O or EtOH as needles. The *hydrochloride* recrystallises from H<sub>2</sub>O or aqueous HCl with m 293°dec (also 298°dec, 300°dec). [Wibaut *Rec Trav Chim Pays Bas* 63 141 1944, IR: Zacharius et al. *J Am Chem Soc* 76 2908 1954, *Beilstein* 22/1 V 244.]

**Polypeptides.** These are strings of  $\alpha$ -amino acids usually with the natural *S*(L) [L-cysteine is an exception and has the *R* absolute configuration] or sometimes “unnatural” *R*(D) configuration at the  $\alpha$ -carbon atom. They generally have less than ~100 amino acid residues. They can be naturally occurring or, because of their small size, can be synthesised chemically from the desired amino acids. Their properties can be very similar to those of small proteins. Many are commercially available, and can be custom made commercially or locally with a peptide synthesiser. They are purified by HPLC and can be used without further purification. Their purity can be checked as described under proteins (Introduction).

**L-Proline** [147-85-3] M 115.1, m 215-220°(dec)(D-isomer), 220-222°(dec) (L-form), 205°(dec)(DL-isomer),  $[\alpha]_D^{20}$  -53° (c 0.6, 0.5N HCl), -93° (c 2.4, 6N KOH) for L-isomer, pl 6.3, pK<sub>1</sub><sup>25</sup> 1.95, pK<sub>2</sub><sup>25</sup> 10.64. A likely impurity is hydroxyproline. Purify L-proline *via* its *picrate* which is crystallised twice from water, then decomposed with 40% H<sub>2</sub>SO<sub>4</sub>. The *picric acid* is extracted with diethyl ether, the H<sub>2</sub>SO<sub>4</sub> in solution is precipitated with Ba(OH)<sub>2</sub>, and the filtrate is evaporated. The residue is crystallised from hot absolute EtOH [Mellan & Hoover *J Am Chem Soc* 73 3879 1951] or EtOH/Et<sub>2</sub>O. Its solubility in H<sub>2</sub>O is >100%. It sublimes

at 182-187°/0.3mm with 99.4% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. It is *hygroscopic* and is stored in a desiccator. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2178-2199 1961, *Beilstein* **22** III/IV 8, **22/1** V 31.]

**L-Prolylglycine** [2578-57-6] **M 172.2, m 236°**,  $[\alpha]_D^{20} +21.1^\circ$  (c 4, H<sub>2</sub>O), **pK<sub>1</sub><sup>25</sup> 3.19, pK<sub>2</sub><sup>25</sup> 8.97**. Recrystallise L-prolylglycine from water at 50-60° by addition of EtOH. [Appel et al. *Chem Ber* **108** 2680 1975, Rydon & Smith *J Chem Soc* 3642 1956.]

**L-Propargylglycine (S-2-aminopent-4-ynoic acid)** [23235-01-0] **M 113.1, m 230°(dec starting at 210°)**,  $[\alpha]_D^{20} -35^\circ$  (c 1, H<sub>2</sub>O),  $-4^\circ$  (c 5, 5N HCl), **pK<sub>Est(1)</sub>~ 2.3 (CO<sub>2</sub>H), pK<sub>Est(2)</sub>~ 9.8 (NH<sub>2</sub>)**. The acid crystallises readily when ~4g in 50ml H<sub>2</sub>O are treated with absolute EtOH at 4°/3 hours, and is collected, washed with cold absolute EtOH and Et<sub>2</sub>O and dried in a vacuum. Also, it recrystallises from aqueous Me<sub>2</sub>CO, R<sub>F</sub> on SiO<sub>2</sub> TLC plates with *n*-BuOH/H<sub>2</sub>O/AcOH (4:1:1) is 0.26. The *racemate* has **m 238-240°**. [Leukart et al. *Helv Chim Acta* **59** 2181 1976, Eberle & Zeller *Helv Chim Acta* **68** 1880 1985, Jansen et al. *Rec Trav Chim Pays-Bas* **88** 819 1969.] It is a suicide inhibitor of  $\gamma$ -cystathionase and other enzymes [Washtier & Abeles *Biochemistry* **16** 2485 1977, Shinozuka et al. *Eur J Biochem* **124** 377 1982].

**R-Pyroglutamic acid (5-oxo-D-proline, R-2-pyrrolidone-5-carboxylic acid)** [4042-36-8] **M 129.1, m 156-158°**,  $[\alpha]_D^{20} +11.2^\circ$  (c 1, H<sub>2</sub>O). Purify *R*-pyroglutamic acid by dissolving it in H<sub>2</sub>O, filtering, passing the filtrate through Dowex 50 (H<sup>+</sup> form), washing with H<sub>2</sub>O, pooling washings, evaporating, removing H<sub>2</sub>O azeotropically with Me<sub>2</sub>CO and \*C<sub>6</sub>H<sub>6</sub>, washing the residue with Et<sub>2</sub>O and recrystallising from EtOH/petroleum ether. [Pradeller et al. *Col Czech Chem Commun* **42** 79, 80 1977, *Beilstein* **22/6** V 7.]

**S-Pyroglutamic acid (5-oxo-L-proline)** [98-79-3] **M 129.1, m 156-158°, 162-164°**,  $[\alpha]_{546}^{20} -11^\circ$  (c 5, H<sub>2</sub>O), **pK<sup>25</sup> 12.7 (by electron spin resonance)**. Crystallise *S*-pyroglutamic acid by dissolving it in boiling EtOH (20g in 100ml), cooling and after a few minutes adding petroleum ether (b 40-60°, 120ml), then after 5 minutes adding a further 120ml, and cooling to room temperature with 90% recovery. This has **m 155.5-157.5°** and  $[\alpha]_D^{20} -11.4^\circ$  (c 4.4, H<sub>2</sub>O) [Hardy *Synthesis* 290 1978, Pellegata et al. *Synthesis* 614 1978]. The NH<sub>4</sub> salt has **m 184-186°** (from EtOH). [*Beilstein* **22/6** V 7.]

**Quisqualic acid (3-[3,5-dioxo-1,2,4-oxadiazolin-2-yl]-L-alanine)** [52809-07-1] **M 189.1, m 190-191°**,  $[\alpha]_D^{20} +17^\circ$  (c 2, 6M HCl), **pK<sub>Est(1)</sub>~ 2.1 (CO<sub>2</sub>H), pK<sub>Est(2)</sub>~ 8.9 (NH<sub>2</sub>)**. It has been purified by ion-exchange chromatography on Dowex 50W (x 8, H<sup>+</sup> form); the desired fractions are lyophilised and recrystallised from H<sub>2</sub>O/EtOH. It has IR (KBr) with  $\nu_{\max}$  at 3400-2750br, 1830s, 1775s, 1745s and 1605s cm<sup>-1</sup>; and <sup>1</sup>H NMR (NaOD/D<sub>2</sub>O, pH 13)  $\delta$ : 3.55-3.57 (1H m, X of ABX, H-2), 3.72-3.85 (2H, AB of ABX, H-3), <sup>13</sup>C NMR (D<sub>2</sub>O)  $\delta$ : 50.1(t), 53.4(d), 154.8(s), 159.7(s) and 171.3(s). [Baldwin et al. *J C S, Chem Commun* 256 1985.] It is a quisqualate receptor agonist [Joels et al. *Proc Natl Acad Sci USA* **86** 3404 1989].

**Sarcosine (N-methylglycine)** [107-97-1] **M 89.1, m 2 12-213°(dec)**, **pK<sub>1</sub><sup>20</sup> 2.12, pK<sub>2</sub><sup>20</sup> 10.19**. Crystallise sarcosine from absolute EtOH, 95% EtOH or H<sub>2</sub>O. It sublimes at 180-185°/0.3mm with 99.1% recovery [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. [Cocker & Harris *J Chem Soc* 1291 1940, Cocker & Lapworth *J Chem Soc* 1897 1931, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** p 2750 1961, *Beilstein* **4** III 1121, **4** IV 2363.]

**Sarcosine anhydride (1,4-dimethylpiperazin-2,5-dione)** [5076-82-4] **M 142.2, m 146-147°, 148°, pK<sub>Est(1)</sub>~ -4.2, pK<sub>Est(2)</sub>~ -1.9**. Crystallise the anhydride from H<sub>2</sub>O, EtOH or EtOAc. Dry it in a vacuum at room temperature. [Karrer et al. *Helv Chim Acta* **5** 140 1922, *Beilstein* **24** II 144, **24** IV 1072.]

**Seleno-DL-methionine ( $\pm$ 2-amino-4-methylselenylbutyric acid)** [1464-42-2, 2578-28-1 ( $\pm$ )] **M 196.1, m 265°(dec), 267-269°(dec), 270° (see pKs of methionine)**. It crystallises in hexagonal plates from MeOH and H<sub>2</sub>O. [Klosterman & Painter *J Am Chem Soc* **69** 2009 1949.] The *L*-isomer [3211-76-5] is purified by dissolving it in H<sub>2</sub>O, adjusting the pH to 5.5 with aqueous NH<sub>3</sub>, evaporating to near-dryness, and the residue is

washed several times with absolute EtOH till a solid is formed and then recrystallise from Me<sub>2</sub>CO. It has *m* 266-268°(dec) [also 275°(dec)], and  $[\alpha]_D^{25} +18.1^\circ$  (c 1, N HCl). [Pande et al. *J Org Chem* **35** 1440 1970, *Beilstein* **4** IV 3216.]

**L-Serine** [56-45-1] *M* 105.1, *m* 228°(dec), 233-235°(dec),  $[\alpha]_D^{25} +14.5^\circ$  (1M HCl),  $[\alpha]_{546}^{20} +16^\circ$  (c 5, 5M HCl), *pK*<sub>1</sub><sup>25</sup> 2.15, *pK*<sub>2</sub><sup>25</sup> 9.21. A likely impurity is glycine. Crystallise L-serine from H<sub>2</sub>O by adding 4 volumes of EtOH. Dry and store it in a desiccator. It sublimes at 160-170°/0.3mm with 99.7% recovery, and unracemised [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. [Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 3** pp 2202-2235 1961, *Beilstein* **4** IV 3118.]

**Somatostatin** [38916-34-6] *M* 1637.9,  $[\alpha]_D^{25} -36^\circ$  (c 0.57, 1% AcOH). Somatostatin is a tetradecapeptide which is purified by gel filtration on Sephadex G-25, eluting with 2N AcOH, and then by liquid partition chromatography on Sephadex G-25 using *n*-BuOH/AcOH/H<sub>2</sub>O (4:1:5) and has *R*<sub>F</sub> = 0.4. It is a brain growth hormone releasing-inhibiting factor which has also been synthesised. [Burgus et al. *Proc Natl Acad Sci USA* **70** 684 1973, Sorantakis & McKinley *Biochem Biophys Res Commun* **54** 234 1973, Hartridt et al. *Pharmazie* **37** 403 1982.]

**L-Threonine** [72-19-5] *M* 119.1, *m* 251-253°, 254°(dec), 262-263°(dec),  $[\alpha]_D^{26} -28.4^\circ$  (H<sub>2</sub>O), *pK*<sub>1</sub><sup>25</sup> 2.17, *pK*<sub>2</sub><sup>25</sup> 9.00. Likely impurities are *allo*-threonine and glycine. Crystallise L-threonine from H<sub>2</sub>O by adding 4 volumes of EtOH. Dry and store it in a desiccator. It also crystallises from 80% EtOH to give hexagonal plates *m* 262-263°(dec). It sublimes at 200-226°/0.3mm with 99.6% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* **77** 1678 1955]. [Elliot *J Chem Soc* **62** 1950, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, **Vol 1** pp 176-183, **Vol 3** pp 2238-2257 1961, *Beilstein* **4** IV 3171.]

**L-Thyroxine sodium salt (5H<sub>2</sub>O)** [6106-07-6] *M* 888.9,  $[\alpha]_{546}^{20} +20^\circ$  (c 2, 1M HCl + EtOH, 1:4). Crystallise the sodium salt from absolute EtOH and dry it for 8 hours at 30°/1mm. [Canepa *Acta Cryst* **4** 283 1951, *Beilstein* **14** II 378, **14** III 1566, **14** IV 2374.]

**D-Thyroxine** {*O*-[3,5-diiodo-4-oxophenyl]-3,5-diiodo-D(-)-tyrosine, 3,3',5,5'-tetraiodo-D-thyronine} [51-49-0] *M* 776.9, *m* 235°(dec), 235-236°(dec), 340°(dec),  $[\alpha]_D^{20} +4.5^\circ$  (c 3, aqueous 0.2N NaOH in 70% EtOH),  $[\alpha]_D^{20} -17^\circ$  (c 2, aqueous N HCl + EtOH 1:4), *pK*<sub>1</sub><sup>25</sup> 2.2 (CO<sub>2</sub>H), *pK*<sub>2</sub><sup>25</sup> 8.40 (OH), *pK*<sub>3</sub><sup>25</sup> 10.1 (NH<sub>2</sub>). Recrystallise D-thyroxine from H<sub>2</sub>O (needles) or from an ammonical solution by dilution with H<sub>2</sub>O, MeOH or Me<sub>2</sub>CO. It has also been purified by dissolving ~6.5 g in a mixture of MeOH (200ml) and 2N HCl (20ml), adding charcoal, filtering then adding NaOAc solution to pH 6. On standing the thyroxine separates, it is filtered off, washed with MeOH then Me<sub>2</sub>CO and dried *in vacuo*. *N*-Formyl-D-thyroxine has *m* 210° and  $[\alpha]_{546}^{21} -26.9^\circ$  (c 5, EtOH). (±)-Thyroxine has *m* 256° and is purified in the same way. [Nahm & Siedel *Chem Ber* **96** 1 1963, Salter *Biochem J* **24** 471 1930, *Beilstein* **14** I 671, **14** II 384, **14** III 1566, **14** IV 2374.]

**L-Thyroxine** (*O*-[3,5-diiodo-4-oxophenyl]-3,5-diiodo-L-(+)-tyrosine, 3,3',5,5'-tetraiodo-L-thyronine) [51-48-9] *M* 776.9, *m* 229-230°(dec), ~235°(dec), 237°(dec),  $[\alpha]_D^{22} -5.1^\circ$  (c 2, aqueous N NaOH + EtOH 1:2),  $[\alpha]_D^{22} +15^\circ$  (c 5, aqueous N HCl in 95% EtOH 1:2),  $[\alpha]_D^{25} +26^\circ$  (EtOH/1M aqueous HCl, 1:1) (*pK*<sub>1</sub><sup>25</sup> 6.6). Purification is the same as for the D-isomer above. Likely impurities are tyrosine, iodotyrosine, iodothyroxines and iodide. Dissolve it in dilute ammonia at room temperature, then crystallise it by adding dilute acetic acid to pH 6. *N*-Formyl-L-thyroxine has *m* 214°(dec) and  $[\alpha]_{546}^{21} +27.8^\circ$  (c 5, EtOH). [Harrington et al. *Biochem J* **39** 164 1945, Nahm & Siedel *Chem Ber* **96** 1 1963, Reineke & Turner *J Biol Chem* **161** 613 1945, Chalmers et al. *J Chem Soc* 3424 1949, *Beilstein* **14** II 378, **14** III 1566, **14** IV 2373.]

***N*-Tosyl-L-lysine chloromethyl ketone (3*S*-1-chloro-3-tosylamino-7-amino-2-heptanone HCl)** [4272-74-6] *M* 369.3, *m* 150-153°(dec), 156-158°(dec), ~165°(dec),  $[\alpha]_D^{20} -7.3^\circ$  (c 2, H<sub>2</sub>O), *pK*<sub>Est</sub> ~ 10.6 (7-NH<sub>2</sub>). The hydrochloride slowly crystallises from a concentrated solution in absolute EtOH, thinned with EtOH/Et<sub>2</sub>O for collection and dried *in vacuo*. It is a suicide enzyme inhibitor of serine proteases, e.g. trypsin and clostripain. [Matsuda et al. *Chem Pharm Bull Jpn* **30** 2512 1982, Shaw et al. *Biochemistry* **4** 2219 1965].

**Triglycyl glycine (tetraglycine)** [637-84-3] M 246.2, m 270-275°(dec),  $pK_1^{25}$  3.21(CO<sub>2</sub>H),  $pK_2^{25}$  7.94(NH<sub>3</sub><sup>+</sup>). Crystallise it from H<sub>2</sub>O (optionally, by the addition of EtOH). [Li et al. *J Am Chem Soc* 79 5859 1957, Rising et al. *J Am Chem Soc* 56 1179 1934, *Beilstein* 4 II 807, 4 III 1201, 4 IV 2472.]

**Trigonellamide chloride (1-methylnicotinamide chloride)** [1005-24-9] M 172.6, m 240°(dec). It crystallises from MeOH, and is dried *in vacuo*. It is prepared from nicotinamide and MeI in refluxing MeOH then shaking with AgCl [Karrer et al. *Helv Chim Acta* 19 826 1936]. It is soluble in organic solvents but moderately in H<sub>2</sub>O. It is a metabolite of nicotinic acid in man, and was isolated from urine [Huff & Perizweig *J Biol Chem* 150 395 1943]. With ketones in aqueous alkali, it produces a green-blue fluorescence which turns blue on acidification and intensifies on heating. [*Beilstein* 22 III/IV 468, 22/2 V 80.]

**3,3',5-Triiodo-S-thyronine** [6893-02-3] M 651.0, m 236-237°(dec),  $[\alpha]_D^{29.5}$  +21.5° (EtOH/1M aqueous HCl, 2:1),  $pK_1^{25}$  6.48,  $pK_2^{25}$  7.62,  $pK_3^{25}$  7.82. Likely impurities are as in *thyroxine*. Purify it by dissolving in dilute NH<sub>3</sub> at ~20°, then crystallise it by addition of dilute acetic acid to pH 6. Alternatively, 35g are purified by dissolving it in a mixture of EtOH (250ml) and 2N NaOH (100ml), then hot 2N HCl is added to the boiling solution until the pH is 4-5. After cooling for a few hours, the solid is filtered off and dried in a vacuum [m 233-235°(dec)]. [Chambers et al. *J Chem Soc* 2433 1949, *Beilstein* 14 III 1566, 14 IV 2373.]

**N,N,N-Trimethyl glycinehydrazide chloride (Girard Reagent T, 2-hydrazino-N,N,N-trimethyl-2-oxoethanaminium chloride)** [123-46-6] M 167.6, m 192°. It is purified by crystallisation from absolute EtOH (slight decomposition) until it has only a slight odour. Store it in well-stoppered containers because it is very hygroscopic. It is very soluble in H<sub>2</sub>O, AcOH and glycerol but slightly soluble in EtOH (0.66%). It forms water-soluble hydrazones with carbonyl compounds. [*Beilstein* 4 III 1133.]

**N-Tris-(hydroxymethyl)methylglycine (TRICINE)** [5704-04-1] M 179.2, m 186-188°(dec),  $pK_1^{20}$  ~2.3,  $pK_2^{20}$  8.15. Crystallise Tricine from EtOH and water. [Good et al. *Methods Enzymol* 24B 53 1968, McGothlin & Jordan *Analyt Lett* 9 245 1976, *Beilstein* 18 III/IV 3454.]

**L-Tryptophan** [73-22-3] M 204.3, m 278°, 281-282°, 290°,  $[\alpha]_D^{20}$  -33.4° (EtOH),  $[\alpha]_{546}^{20}$  -36° (c 1, H<sub>2</sub>O),  $[\alpha]_D^{28}$  +28° (c 2.1, 10% HCl),  $pK_1^{25}$ -6.23 (aqueous H<sub>2</sub>SO<sub>4</sub>),  $pK_2^{25}$  2.46,  $pK_3^{25}$  9.41,  $pK_4^{25}$  14.82 (acidic NH, in aqueous NaOH). Crystallise L-tryptophan from H<sub>2</sub>O/EtOH, wash it with anhydrous diethyl ether and dry it at room temperature in a vacuum over P<sub>2</sub>O<sub>5</sub>. It sublimes at 220-230°/0.03mm with 99% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Cox & King *Org Synth Coll Vol II* 612 1943, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2316-2345 1961, *Beilstein* 22 IV 6765.]

**Tyrocidine A (cyclic decapeptide antibiotic with two D-Phe amino acids)** [1481-70-5] M 1268.8, m 240°(dec),  $[\alpha]_D^{25}$  -115° (c 0.91, MeOH). Crystallise tyrocidine A as the *hydrochloride* from MeOH or EtOH/HCl. [Paladin & Craig *J Am Chem Soc* 76 688 1954, King & Craig *J Am Chem Soc* 77 6624 1955, Okamoto et al. *Bull Chem Soc Jpn* 50 231 1977, *Beilstein* 26 III/IV 4280.]

**L-Tyrosine** [60-18-4] M 181.2, m 290-295°(dec), 294-300°(dec),  $[\alpha]_D^{25}$  -10.0° (5M HCl),  $pK_1^{25}$  2.18 (CO<sub>2</sub>H),  $pK_2^{25}$  9.21 (OH),  $pK_3^{25}$  10.47 (NH<sub>3</sub><sup>+</sup>). Likely impurities are L-cysteine and the ammonium salt. L-Tyrosine is dissolved in dilute ammonia, then crystallised by adding dilute acetic acid to pH 5. Also, crystallise it from H<sub>2</sub>O or EtOH/H<sub>2</sub>O, and dry it at room temperature in a vacuum over P<sub>2</sub>O<sub>5</sub>. It sublimes at 235-240°/0.03mm with 99.2% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Albert *Biochem J* 50 690 1952, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2348-2366 1961, *Beilstein* 14 IV 2264.]

**L-Valine** [72-18-4] M 117.2, m 305-308°(dec), 315°,  $[\alpha]_D^{17}$  +28.4° (c 2, 5M HCl),  $pK_1^{20}$  2.38 (CO<sub>2</sub>H),  $pK_2^{20}$  9.59 (NH<sub>3</sub><sup>+</sup>). Crystallise L-valine from water by addition of EtOH. It sublimes at 178-188°/0.03mm with 99.3% recovery and unracemised [Gross & Gradsky *J Am Chem Soc* 77 1678 1955]. [Perrin *J Chem Soc* 3125 1958, Greenstein & Winitz *The Chemistry of the Amino Acids* J. Wiley, Vol 3 pp 2368-2377 1961, *Beilstein* 4 IV 2659.]

## PROTEINS, ENZYMES, DNA and RNA

**Abrin A and C (agglutinins from *Abrus seeds*)** [1393-62-0]  $M_r$  63,000-67,000. These are toxic lectins (proteins) from seeds of *Abrus precatorius*. The yellow-white powder is purified by successive chromatography on DEAE-Sephadex A-50, carboxymethylcellulose, and DEAE-cellulose. Abrin A is more positively charged on the DEAE-cellulose column and has been crystallised from  $(\text{NH}_4)_2\text{SO}_4$  by the free interface diffusion technique. Its molecular weight (by sedimentation equilibrium) is 60,000, whereas Abrin C has molecular weight of 63,800. Treatment of A with mercaptoethanol at 100°/2 hours followed by SDS-PAGE gave a main band with  $M_r$  32,000 and two very weak bands, whereas C (which is more toxic) gave two intense bands with  $M_r$  28,000 and 33,000. [Wei et al. *J Biol Chem* **249** 3061 1974.] Abrin C has been crystallised for X-ray analysis by the free interface diffusion technique described by Salemm [Arch Biochim Biophys **151** 533 1972]. The crystals were grown at 37° in Pyrex tubes (5 x 30 cm) by layering 50 $\mu$ l of protein solution (22mg/ml) over 100 $\mu$ l of unbuffered 70% saturated  $(\text{NH}_4)_2\text{SO}_4$  [Wei & Einstein *J Biol Chem* **249** 2985 1974.] [UV and CD: Herrmann & Behnke *Biochem Biophys Acta* **621** 43 1980, for physical and chemical properties see *Biochem Biophys Acta* **667** 397 1981, Beilstein **22** III/IV 6776.]

**Acetoin dehydrogenase [from beef liver, acetoin NAD oxidoreductase]** [9028-49-3]  $M_r$  76000 [EC 1.1.1.5]. Purify it via the acetone cake, then Ca-phosphate gel filtration (unabsorbed), lyophilised and then fractionated through a DEAE-22 cellulose column. The  $K_m$  for diacetyl is 40 $\mu$ M, and for NADH it is 100 $\mu$ M in phosphate buffer at pH 6.1. [Burgos & Martin *Biochim Biophys Acta* **268** 261 1972, **289** 13 1972.]

**$\beta$ -D-N-Acetylglucosaminidase [from *M sexta* insects]** [9012-33-3]  $M_r$  ~61,000 [EC 3.2.1.52]. Purify it by chromatography on DEAE-Biogel, hydroxylapatite chromatography and gel filtration through Sephacryl S200. Two isoforms: a hexosaminidase EI with  $K_m$  177 $\mu$ M ( $V_{max}$  328 sec<sup>-1</sup>) and EII a chitinase with  $K_m$  160 $\mu$ M ( $V_{max}$  103 sec<sup>-1</sup>) with 4-nitrophenyl- $\beta$ -acetylglucosamine as substrate. [Dziadil-Turner *Arch Biochim Biophys* **112** 546 1981.]

**$\beta$ -D-N-Acetylhexosaminidase A and B (from human placenta)** [9012-33-3]  $M_r$  ~61,000 [EC 3.2.1.52]. Purify it by Sephadex G-200 filtration and DEAE-cellulose column chromatography. The hexosaminidase A is further purified by DEAE-cellulose column chromatography, followed by an ECTEOLA-cellulose column, Sephadex-200 filtration, electrofocusing and Sephadex G-200 filtration. Hexosaminidase B is purified by a CM-cellulose column, electrofocusing and Sephadex G-200 filtration. [Srivastava et al. *J Biol Chem* **249** 2034 1974.]

**N-Acetyl neuraminic acid aldolase [from *Clostridium perfringens*, N-acetylneuraminic acid pyruvate lyase]** [9027-60-5]  $M_r$  32,000 [EC 4.1.3.3]. Purify the aldolase by extraction with H<sub>2</sub>O, protamine precipitation,  $(\text{NH}_4)_2\text{SO}_4$  fractionation, Me<sub>2</sub>CO precipitation, acid treatment at pH 5.7 and precipitation at pH 4.5. The equilibrium constant for pyruvate + *n*-acetyl-D-mannosamine  $\rightleftharpoons$  N-acetylneuraminidate at 37° is 0.64. The  $K_m$  for N-acetylneuraminic acid is 3.9mM in phosphate at pH 7.2 and 37°. [Comb & Roseman *Methods Enzymol* **5** 391 1962.] The enzyme from hog kidney (cortex) has been purified 1700-fold by extraction with H<sub>2</sub>O, protamine sulfate precipitation,  $(\text{NH}_4)_2\text{SO}_4$  fractionation, heating between 60-80°, a second  $(\text{NH}_4)_2\text{SO}_4$  fractionation and starch gel electrophoresis. The  $K_m$  for N-acetylneuraminic acid is 1.5mM. [Brunetti et al. *J Biol Chem* **237** 2447 1962.]

**Acyl-coenzyme A Synthase [from beef liver]** [9013-18-7]  $M_r$  57,000 [EC 6.2.1.2]. Purify the synthase by extraction with sucrose/HCO<sub>3</sub> buffer, protamine sulfate precipitation,  $(\text{NH}_4)_2\text{SO}_4$  (66-65%) fractionation (pH 4.35) and a second  $(\text{NH}_4)_2\text{SO}_4$  (35-60%) fractionation (pH 4.35). It has  $K_m$  0.15mM ( $V_{rel}$  1.0) for octanoate and 0.41mM ( $V_{rel}$  2.37) for heptanoate. The  $K_m$  for ATP is 0.5mM, all at pH 9.0 in ethylene glycol buffer at 38°. [Jencks et al. *J Biol Chem* **204** 453 1953, *Methods Enzymol* **5** 467 1962.]

**Acyl-coenzyme A Synthase (from yeast)** [9012-31-1]  $M_r$  ~151,000 [EC 6.2.1.1]. This enzyme has been purified by extraction into phosphate buffer pH 6.8-7.0 containing 2-mercaptoethanol and EDTA, protamine sulfate precipitation, polyethylene glycol fractionation, Alumina  $\gamma$  gel filtration, concentration by  $(\text{NH}_4)_2\text{SO}_4$

precipitation, Bio-Gel A-0.5m chromatography and DEAE-cellulose gradient chromatography. It has  $K_m$  (apparent) 0.24mM (for acetate) and 0.035mM (for CoA); 1.2 mM for ATP and  $Mg^{2+}$  4.0mM. [Frenkel & Kitchens *Methods Enzymol* **71** 317 1981.]

**ADP-Ribosyl transferase (adenyl transferase, polynucleotide, from human placenta)** [9026-30-6]  $M_r \sim 115,000$  [EC 2.4.2.30]. Purify the transferase by making an affinity absorbent for ADP-ribosyltransferase by coupling 3-aminobenzamide to Sepharose 4B. [Burtscher et al. *Anal Biochem* **152** 285 1986.]

**Agglutinin (from peanuts)** [*Arachis hypogaea*] [1393-62-0]  $M_r$  **134,900 (tetramer)**. Agglutinin is purified by affinity chromatography on Sepharose- $\zeta$ -aminocaproyl- $\beta$ -D-galactopyranosylamine. [Lotan et al. *J Biol Chem* **250** 8518 1974.]

**Albumin (bovine and human serum)** [9048-46-8 (bovine), 70024-90-7 (human)]  $M_r \sim 67,000$  (bovine), **69 000 (human)**, UV:  $A_{280nm}^{1\%}$  **6.6 (bovine) and 5.3 (human) in H<sub>2</sub>O**,  $[\alpha]_{546}^{25}$  **-78.2° (H<sub>2</sub>O)**. Albumin is purified by dissolving it in conductivity water and passage at 2-4° through two ion-exchange columns, each containing a 2:1 mixture of anionic and cationic resins (Amberlite IR-120, H-form, Amberlite IRA-400, OH<sup>-</sup> form). This treatment removes ions and lipid impurities. Care is taken to exclude CO<sub>2</sub>, and the solution is stored at -15°. [Möller et al. *Trans Faraday Soc* **57** 312 1961.] More complete lipid removal is achieved by lyophilising the deionised solution, covering the dried albumin (human serum) with a mixture of 5% glacial acetic acid (v/v) in iso-octane (previously dried with Na<sub>2</sub>SO<sub>4</sub>) and allowing it to stand at 0° (without agitation) for upwards of 6 hours before decanting and discarding the extraction mixture, washing with iso-octane, re-extracting, and finally washing twice with iso-octane. The purified albumin is dried under vacuum for several hours, then dialyzed against water for 12-24 hours at room temperature, lyophilised, and stored at -10°C [Goodman *Science* **125** 1296 1957]. It has been recrystallised in high (35%) and in low (22%) EtOH solutions from Cohn's Fraction V.

The **high EtOH recrystallisation** is as follows: To 1kg of Fraction V albumin paste at -5° is added 300ml of 0.4 M pH (pH 5.5) acetate buffer in 35% EtOH pre-cooled to -10° and 430 ml of 0.1 M NaOAc in 25% EtOH also at -10°. Best results are obtained by adding all of the buffer and about half of the NaOAc and stirring slowly for 1 hour. The rest of the NaOAc is added when all the lumps have disintegrated. The mixture is set aside at -5° for several days to crystallise. 35% EtOH (1 L) is then added to dilute the crystalline suspension and lower the ionic strength prior to centrifugation at -5° (yield 80%). The crystals are further dissolved in 1.5 volumes of 15% EtOH/0.02M NaCl at -5° and clarified by filtration through washed, calcined diatomaceous earth. This solution may be recrystallised by re-adjusting to the conditions in the first crystallisation, or it may be recrystallised at 22% EtOH with the aid of a very small amount of decanol (enough to give a final concentration of 0.02%). Note that crystallisation from lower EtOH concentration gave better purification (i.e. by removing globulins and carbohydrates) and producing a more stable product.

The **low EtOH recrystallisation** is as follows: To 1kg of Fraction V at -10° to -15° is added 500ml of 15% EtOH at -5°, stirred slowly until a uniform suspension is formed. To the 15% EtOH (500ml) is added sufficient 0.2M NaHCO<sub>3</sub> solution (125-150ml) at 0° to bring the pH (1:10 dilution) to 5.3. Some temperature rise occurs, and care must be taken to keep the temperature < -5°. If the albumin is incompletely dissolved a small amount of H<sub>2</sub>O is added (100ml at a time at 0°, allowing 15 minutes between additions). Undissolved albumin can be easily distinguished from small amounts of undissolved globulins, or as the last albumin dissolves, the appearance of the solution changes from milky white to hazy grey-green in colour. Keep the solution at -5° for 12 hours and filter by suspending in 15g of washed fine calcined diatomaceous earth, and filtering using a Büchner funnel precoated with coarser diatomaceous earth. The filtrate may require two or more similar filtrations to give a clear solution. To crystallise the filtrate, add through a capillary pipette, and with careful stirring, 1/100 volume of a solution containing 10% decanol and 60% EtOH (at -10°), and seed with the needle-type albumin crystals. After 2-3 days, crystallisation is complete. The crystals are centrifuged off. These are suspended with gentle mechanical stirring in one-third their weight of 0.005 M NaCl pre-cooled to 0°. With careful stirring, H<sub>2</sub>O (at 0°) is added slowly in an amount equal to 1.7 times the weight of the crystals. At this stage there is about 7% EtOH, and the temperature cannot be made lower than -2.5° to -1°. Clarify, and collect as above. [Cohn et al. *J Am Chem Soc* **69** 1753 1947.]

Human serum albumin has been purified similarly with 25% EtOH and 0.2% decanol. The isoelectric points of bovine and human serum albumins are 5.1 and 4.9, respectively.

**Angiotensin (from rat brain)** [70937-97-2]  $M_r$  1524.8. Angiotensin is purified using extraction, affinity chromatography and HPLC [Hermann et al. *Anal Biochem* **159** 295 1986].

**Angiotensin-converting enzyme (ACE, peptidyl peptide hydrolase) (from rabbit lung)** [9015-82-1]  $M$  129,000 Dal (equilibrium sedimentation),  $M_r$  ~140,000 (SDS-PAGE) [EC 3.14.15.1]. Purify ACE by fractionation on DEAE-cellulose, Ca phosphate gel chromatography, elution from Sephadex G-200 and lectin affinity chromatography. The MW varied with glycosidation and is higher by gel filtration. It contains one atom of Zn/mol and has  $K_m$  values for hydrolysis of hippurylhistidinylleucine and angiotensin I of 2.3 and 0.07 mM, and turnover of 15,430 and 792 mol/min/mol at 37°, respectively. The activity is inhibited by EDTA and increased amounts of Ca ions but required Ca ions. [Das & Soffer *J Biol Chem* **250** 6762 1975, Reviewed by Ehlers & Riordan *Biochemistry* **28** 5311 1989.]

**Angiotensinogen (from porcine plasma)** [64315-16-8]  $M_r$  59,400, 60,000, 62,600, 63,600 depending on sialic acid content. This rennin substrate is purified 390-fold from the serum by chromatography on Blue Sepharose, phenyl sepharose, hydroxyapatite and finally by affinity chromatography on 5-hydroxytryptamine (5-HT) sepharose to which it specifically binds to the 5-HT. It is applied to the latter column in 50mM sodium phosphate at pH 7 and after washing, it is eluted by increasing the ionic strength with 100mM sodium phosphate buffer containing 250mM NaCl. The multiple forms are separated by SDS-PAGE and have pI 4.40-4.82. [Campbell et al. *Biochem J* **243** 121 1987.]

**Avidin (from egg white)** [1405-69-2]  $M_r$  ~70,000. Avidin is purified by chromatography of an ammonium acetate solution on CM-cellulose [Green *Biochem J* **101** 774 1966]. It is also purified by affinity chromatography on 2-iminobiotin-6-aminoethyl-Sepharose 4B [Orr *J Biol Chem* **256** 761 1981]. It is a biotin-binding protein.

**Azurin (from *Pseudomonas aeruginosa*)** [12284-43-4]  $M_r$  30,000. Azurin with  $A_{625/280} = 0.56$  is purified by gel chromatography on G-25 Sephadex with 5mM phosphate pH 7 buffer as eluent [Cho et al. *J Phys Chem* **91** 3690 1987]. It is a blue Cu protein used in biological electron transport, and its reduced form is obtained by adding a slight excess of  $Na_2S_2O_4$ . [See *Structure and Bonding* Springer Verlag, Berlin **23** 1 1975.]

**Bromelain (anti-inflammatory Ananase from pineapple)** [37189-34-7]  $M_r$  ~33,000 [EC 3.4.33.4]. This protease has been purified *via* the acetone powder, G-75 Sephadex gel filtration and Bio-Rex 70 ion-exchange chromatography, and has  $A_{1cm}^{1\%}$  20.1 at 280nm. The protease from pineapple hydrolyses benzoyl glycine ethyl ester with a  $K_m$  (app) of 210mM and  $k_{cat}$  of 0.36 sec<sup>-1</sup>. [Murachi *Methods Enzymol* **19** 273 1970, Balls et al. *Ind Eng Chem* **33** 950 1941.]

**Carbonic anhydrase (carbonate hydrolase)** [9001-03-0]  $M_r$  31,000 [EC 4.2.1.1]. Purify carbonic anhydrase by hydroxylapatite and DEAE-cellulose chromatography [Tiselius et al. *Arch Biochem Biophys* **65** 132 1956, *Biochim Biophys Acta* **39** 218 1960], and is then dialysed for crystallisation. A 0.5 to 1% solution of the enzyme in 0.05 M Tris-HCl pH 8.5 is dialysed against 1.75M solution of  $(NH_4)_2SO_4$  in the same buffer, and this solution is slowly increased in salt concentration by periodic removal of small amounts of dialysate and replacing with an equal volume of 3.5M  $(NH_4)_2SO_4$ . The final salt concentration, in which the DEAE-cellulose fractions give beautiful birefringent suspensions of crystals, ranged from 2.4 to 2.7M and appeared first as fine crystals, then underwent transition to thin fragile plates. Carbonic anhydrase is a Zn enzyme which exists as several isoenzymes of varying degrees of activity [J *Biol Chem* **243** 6474 1968, crystal structure: *Nature, New Biology* **235** 131 1972; see also P.D. Boyer Ed. *The Enzymes* Academic Press NY, pp 587-665 1971].

**Carboxypeptidase A (from bovine pancreas, peptidyl-L-aminoacid lyase)** [11075-17-5]  $M_r$  34,600 [EC 3.4.17.1]. Carboxypeptidase A is purified by DEAE-cellulose chromatography, activation with trypsin and dialysed against 0.1M NaCl, yielding crystals. It is recrystallised by dissolving in 20 ml of M NaCl and dialysed for 24 hours each against the following salts present in 500ml of 0.02M sodium veronal pH 8.0, 0.5M NaCl, 0.2M NaCl and 0.15M NaCl. The last dialysate usually induces crystallisation. If it does not crystallise, then

dialyse the last solution against 0.02M sodium veronal containing 0.10M NaCl. Only 2 or 3 re-crystallisations are required to attain maximum activity. [Cox et al. *Biochemistry* **3** 44 1964.] Enzyme activity is measured by hydrolysing hippuryl-L-phenylalanine (or phenylacetic acid) and observing the rate of change of optical density at 254nm (reaction extinction coefficient is  $\sim 0.592 \text{ cm}^2/\mu\text{mole}$  at pH 7.5) [Bergmyer *Methods in Enzymatic Analysis* (Academic Press) **1** 436 1974].

**Cathepsin B (from human liver)** [9047-22-7]  $M_r$  27,500 [EC 3.4.22.1]. Cathepsin B is purified by affinity chromatography on the semicarbazone of Gly-Phe-glycinal-linked to Sepharose 4B, with elution by 2,2'-dipyridyl disulfide [Rich et al. *Biochem J* **235** 731 1986, *Methods Enzymol* **80** 551 1981].

**Cathepsin D (from bovine spleen)** [9025-26-7]  $M_r$  56,000 [EC 3.4.23.5]. Cathepsin D is purified on a CM column after  $(\text{NH}_4)_2\text{SO}_4$  fractionation and dialysis, then starch-gel electrophoresis and by ultracentrifugal analysis. Finally chromatograph on a DEAE column [Press et al. *Biochem J* **74** 501 1960].

**Ceruloplasmin (from human blood plasma)** [9031-37-2]  $M_r$  134,000. This blue protein is the principal Cu transporter (up to 90% of circulating Cu) and is purified by precipitation with polyethylene glycol 4000, batchwise adsorption and elution from QAE-Sephadex, and gradient elution from DEAE-Sepharose CL-6B. Ceruloplasmin is thus purified 1640-fold and is homogeneous on anionic polyacrylamide gel electrophoresis (PAGE), SDS-PAGE, isoelectric focusing and low-speed equilibrium centrifugation. It has  $\lambda_{\text{max}}$  at 280, 260nm ( $A_{1\text{cm}}^{1\%} \approx 0.68$ ). [Oesthuizen *Anal Biochem* **146** 1 1985, Cohn et al. *J Am Chem Soc* **68** 459 1946.]

**Chemokines.** These are small proteins formed from longer precursors and are chemo-attractants for lymphocytes and lymphoid organs. They are characterised by having cysteine groups in specific relative positions. The two largest families are the  $\alpha$  and  $\beta$  families that have four cysteine residues arranged (C-X-C) and (C-C) respectively. The mature chemokines have  $\sim 70$  amino acids with internal cys S-S bonds and attract myeloid type cells *in vitro*. The  $\gamma$ -family (Lymphotactin) has only two cys residues. The  $\delta$ -family (Neurotactin, Fractalkine) has the C-C-X-X-X-C sequence (*ca* 387 amino acids), binds to membrane and promotes adhesion of lymphocytes. The soluble domain of human Fractalkine "chemo-attracts" monocytes and T cells. Several chemokines are available commercially (some prepared by recombinant DNA techniques), including 6CKine/exodus/SLC which belongs to the  $\beta$ -family with 6 cysteines (110 amino acids, mature protein), as the name implies (C-C-C-C-X.....X-C-C) and homes lymphocytes to secondary lymphoid organs with lymphocyte adhesion antitumor properties. Other chemokines available are C10 ( $\beta$ CC) and Biotaxin. Several chemokine receptors and antibodies are available commercially and can generally be used without further purification. [Murphy "Molecular biology of lymphocyte chemo-attractant receptors" in *Ann Rev Immunol* **12** 593 1994.]

**Chirazymes.** These are commercially available enzymes, e.g. lipases, esterases, that can be used for the preparation of a variety of optically active carboxylic acids, alcohols and amines. They can cause regio and stereospecific hydrolysis and do not require cofactors. Some can be used also for esterification or transesterification in neat organic solvents. The proteases, amidases and oxidases are obtained from bacteria or fungi, whereas esterases are from pig liver and thermophilic bacteria. For preparative work the enzymes are covalently bound to a carrier and do not therefore contaminate the reaction products. Chirazymes are available from Roche Molecular Biochemicals and are used without further purification.

**$\alpha$ -Chymotrypsin** [9004-07-3]  $M_r$   $\sim 25000$  [EC 3.4.21.1].  $\alpha$ -Chymotrypsin is crystallised twice from four-tenths saturated ammonium sulfate solution, then dissolved in 1mM HCl and dialysed against 1mM HCl at 2-4°. The solution is stored at 2° [Lang et al. *J Am Chem Soc* **80** 4923 1958].

**Citric acid cycle components (from rat heart mitochondria).** These are resolved by anion-exchange chromatography [LaNoue et al. *J Biol Chem* **245** 102 1970].

**Clostripain** [9028-00-6]  $M_r$   $\sim 55,000$  [EC 3.4.22.8]. Clostripain is isolated from *Clostridium histolyticum* collagenase by extraction in pH 6.7 buffer, followed by hydroxylapatite chromatography with a 0.1-0.2 M phosphate gradient, then Sephadex G-75 gel filtration with 0.05M phosphate pH 6.7, dialysis and a second

hydroxylapatite chromatography (gradient elution with 0.1M  $\rightarrow$  0.3M phosphate, pH 6.7) purification. It has proteinase and esterase activity and is assayed by hydrolysing *N*-benzoyl-L-arginine methyl ester. [Mitchell & Harrington *J Biol Chem* **243** 4683 1968, *Methods Enzymol* **19** 635 1970.]

**Colicin E1 (from *E. coli*)** [11032-88-5].  $M_r$  56,000, pI 9.5. Colicin E1 is purified (8.6-fold to Specific Activity of  $1.5 \times 10^5$  units/mg) from *E. coli* JC411 by salt extraction of extracellular-bound colicin followed by  $(NH_4)_2SO_4$  (40-60% saturation) fractionation and ion-exchange chromatography on a DEAE-Sephadex A 50 column, and then by CM-Sephadex column chromatography [Schwartz & Helinski *J Biol Chem* **246** 6318 1971].

**Collagenase (from human polymorphonuclear leukocytes)** [9001-12-1]  $M_r$  68,000-125,000 [EC 3.4.24.3]. Collagenase is purified by using *N*-ethylmaleimide to activate the enzyme, and wheat germ agglutinin-agarose affinity chromatography [Callaway et al. *Biochemistry* **25** 4757 1986].

**Copper-zinc-superoxide dismutase (from blood cell haemolysis)** [9054-89-1]  $M_r$  ~32,000 [EC 1.15.1.1]. The dismutase is purified by DEAE-Sephadex and copper chelate affinity chromatography. The preparation is homogeneous by SDS-PAGE, by analytical gel filtration chromatography and by isoelectric focusing [Weselake et al. *Anal Biochem* **155** 193 1986, Fridovich *J Biol Chem* **244** 6049 1969].

**Cytochrome  $c_1$  (from horse, beef or fishes' heart, or pigeon breast muscle)** [9007-43-6]  $M_r$  ~13,000. Cytochrome  $c_1$  is purified by chromatography on CM-cellulose (CM-52 Whatman) [Brautigan et al. *Methods Enzymol* **53D** 131 1978]. It has a high PI (isoelectric point) and has been purified further by adsorption onto an acidic cation exchanger, e.g. Amberlite IRC-50 (polycarboxylic) or in ground form Amberlite XE-40 (100-200 mesh) or Decalco-F (aluminium silicate), where the non-cytochrome protein is not adsorbed and is readily removed. The cytochrome is eluted using a solution containing 0.25g ions/L of a univalent cation at pH 4.7 adsorbed onto the  $NH_4^+$  salt of Amberlite IRC-50 at pH 7, washed with  $H_2O$  and then with 0.12M  $NH_4OAc$  to remove non-cytochrome protein. When the cytochrome begins to appear in the eluate, then the  $NH_4OAc$  concentration is increased to 0.25 M. The fractions with  $ca$   $Fe = 0.465-0.467$  are collected, dialysed against  $H_2O$  and adsorbed onto a small IRC-50 column and eluted with 0.5M  $NH_3$ , then dialysed and lyophilised. (A second fraction II can be eluted from the first resin with 0.5M  $NH_3$  but is discarded). [Keilin & Hartree *Biochemical Preparations* **1** 1 1952, Margoliash *Biochemical Preparations* **8** 33 1957.]

*Cytochrome c has been recrystallised as follows:* The above eluate ( $ca$  100ml) is dialysed against  $H_2O$  (10 vols) at 4° for 24 hours (no more), then passed through an XE-40 column (2 x 1 cm above) which is equilibrated with 0.1M  $NH_4OAc$  pH 7.0. The column is washed with 0.1%  $(NH_4)_2SO_4$  pH 8.0, and the dark red resin in the upper part of the column is collected and in 0.1%  $(NH_4)_2SO_4$  pH 8.0 is transferred to another column (7mm diameter) and the cytochrome c is eluted with 5%  $(NH_4)_2SO_4$  pH 8.0. More than 98% of the red colour is collected in a volume of  $ca$  4ml in a weighed centrifuge tube. Add a drop of octanol and 0.43g of  $(NH_4)_2SO_4/g$  of solution. When the salt has dissolved, ascorbic acid (5mg) is added as well as a few drops of 30% aqueous  $NH_3$ , and it is kept at 10° for 10 minutes (turns lighter colour due to reduction). Then add finely powdered  $(NH_4)_2SO_4$  in small portions (stir with a glass rod) until the solution becomes turbid. Stopper the tube tightly, and set aside at 15-25° for 2 days while the cytochrome c separates as fine needles or rosettes. Further  $(NH_4)_2SO_4$  (20mg) are added per ml of suspension and kept in the cold for a few days to complete the crystallisation. The crystals are collected by centrifugation (5000xg), suspended in saturated  $(NH_4)_2SO_4$  (pH 8.0 at 10°), then centrifuged again. For recrystallisation the crystals are dissolved in the least volume of  $H_2O$ , one drop of ammonia and 1 mg of ascorbic acid are added and the above process is repeated. The yield of twice recrystallised cytochrome c from 2Kg of muscle is  $ca$  200 mg, but this varies with the source and freshness of the muscle used. The crystals are stored as a solid after dialysis against 0.08M NaCl or 0.1M sodium buffer and lyophilising, or as a suspension in saturated  $(NH_4)_2SO_4$  at 0°. [Hagihara et al. *Biochemical Preparations* **6** 1 1958.]

*Purity of cytochrome c:* This is checked by the ratio of the absorbance at 500nm (reduced form) to 280nm (oxidised form), i.e.  $\epsilon_{500}/\epsilon_{280}$  should be between 1.1 and 1.28, although values of up to 1.4 have been obtained for pure preparations.

For the preparation of the *reduced form* see Margoliash *Biochemical Preparations* **5** 33 1957 and Yonetani *Biochemical Preparations* **11** 19 1966.

Cytochrome from *Rhodospirillum rubrum* ( $\epsilon_{270}/\epsilon_{551}$  0.967) is purified by chromatography on a column of CM-Whatman cellulose [Paleus & Tuppy *Acta Chem Scand* **13** 641 1959].

**Cytochrome c oxidase (from bovine heart mitochondria).** [9001-16-5]  $M_r$  100,000/haeme [EC 1.9.3.1]. The oxidase is purified by selective solubilisation with Triton X-100 and subsequently with lauryl maltoside, finally by sucrose gradient centrifugation [Li et al. *Biochem J* **242** 417 1978].

It has also been purified by extraction in 0.02 M phosphate buffer (pH 7.4) containing 2% of cholic acid (an inhibitor which stabilises as well as solubilises the enzyme) and fractionated with  $(\text{NH}_4)_2\text{SO}_4$  collecting the 26-33% saturation cut and refractionating again and collecting the 26-33% saturation fraction. The pellet collected at 10,000x g appears as an oily paste. The cholate needs to be removed to activate the enzyme as follows: The precipitate is dissolved in 10ml of 0.1M phosphate buffer pH 7.4, containing 1% of Tween-80 and dialysed against 1L of 0.01 M  $\text{PO}_4$  buffer (pH 7.4) containing 1% of Tween-80 for 10 hours at 0° and aliquoted. The enzyme is stable at 0° for 2 weeks and at -15° for several months. It is assayed for purity (see reference) by oxidation of reduced cytochrome c ( $K_m$  10 $\mu\text{M}$ ). [Yonetani *Biochemical Preparations* **11** 14 1966, *J Biol Chem* **236** 1680 1961.]

**Cytokines** See chemokines, interferons, interleukins.

**Deoxyribonucleic acid (from plasmids).** These are purified by two buoyant density ultracentrifugations using ethidium bromide-CsCl. The ethidium bromide is extracted with  $\text{Et}_2\text{O}$ , and the DNA is dialysed against buffered EDTA and lyophilised. [Marmur & Doty *J Mol Biol* **5** 109 1962, Guerry et al. *J Bacteriol* **116** 1064 1973.] See “Introduction” in this chapter.

**Dermatan sulfate (chondroitin sulfate B from pig skin)**  $M_r$  20,000-36,000 [54328-33-5 (*Na salt*)]. Dermatan sulfate is purified by digestion with papain and hyaluronidase, and fractionation using aqueous  $\text{EtOH}$ . [Gifonelli & Roden *Biochemical Preparations* **12** 1 1968.]

**Dihydrofolate reductase (from *Mycobacterium phlei*)** [9002-03-3]  $M_r$  ~18,000 [EC 1.5.1.3]. Dihydrofolate reductase is purified by  $(\text{NH}_4)_2\text{SO}_4$  precipitation, then fractionation on a Sephadex G-75 column, applied to a Blue Sepharose column and eluted with 1mM dihydrofolate. [Al Rubeai & Dole *Biochem J* **235** 301 1986.]

**Dihydropteridine reductase (from sheep liver)** [9074-11-7]  $M_r$  52,000 [EC 1.6.99.7]. Dihydropteridine reductase is purified by fractionation with ammonium sulfate, dialysed against Tris buffer, adsorbed and eluted from hydroxylapatite gel. It is then run through a DEAE-cellulose column and also subjected to Sephadex G-100 filtration. [Craine et al. *J Biol Chem* **247** 6082 1972.]

**Dihydropteridine reductase (from human liver)** [9074-11-7]  $M_r$  52,000 [EC 1.6.99.7]. Dihydropteridine reductase is purified to homogeneity on a naphthoquinone affinity adsorbent, followed by DEAE-Sephadex and CM-Sephadex chromatography. [Firgaira, Cotton and Danks, *Biochem J* **197** 31 1981.] [For other dihydropteridine reductases see Armarego et al. *Med Res Rev* **4**(3) 267 1984.]

**3,4-Dihydroxyphenylalanine-containing proteins.** Boronate affinity chromatography is used in the selective binding of proteins containing 3,4-dihydroxyphenylalanine to a *m*-phenylboronate agarose column and eluted with 1M  $\text{NH}_4\text{OAc}$  at pH 10. [Hankus et al. *Anal Biochem* **150** 187 1986.]

**Dipeptidyl aminopeptidase (dipeptidyl peptidase IV, from rat brain)** [9031-94-1, 54249-88-6]  $M_r$  87,500 (monomer SDS-PAGE), (88,107 from nucleotide sequence), up to 4000,000 [EC 3.4.14.5]. The aminopeptidase is purified about 2000-fold by column chromatography on CM-cellulose, hydroxylapatite and Gly-Pro AH-Sepharose. [Imai et al. *J Biochem (Tokyo)* **93** 431 1983, Schomburg & Schomburg *Springer Handbook of Enzymes* 2nd Edn vol 6 p 286 2002.]

**DNA (deoxyribonucleic acids).** The essential structures of chromosomes are DNA and contain the genetic “blueprint” in the form of separate genes. They are made up of the four deoxyribonucleic acids (nucleotides): adenylic acid, guanylic acid, cytidylic acid and thymidylic acid (designated A, G, C, T respectively) linked together by their phosphate groups in ester bonds between the 3' and 5' hydroxy groups of the 2'-deoxy-D-ribose moiety of the nucleotides. The chains form a double-stranded spiral (helix) in which the two identical nucleotide

sequences run antiparallel with the heterocyclic bases hydrogen bonded (A...T, G...C) forming the “ladder” between the strands. Short sequences of DNA are available commercially, are commercially custom made or synthesised in a DNA synthesiser and purified by HPLC. Their purity can be checked by restriction enzyme cleavage followed by gel electrophoresis, or directly by gel electrophoresis or analytical HPLC. Commercial DNAs are usually pure enough for direct use but can be further purified using commercially available kits involving binding to silica or other matrices and eluting with Tris buffers. There are now rapid “throughput” techniques for sequencing DNA which are very accurate.

**Dopamine- $\beta$ -hydroxylase (from bovine adrenal medulla)** [9013-38-1]  $M_r \sim 290,000$  [EC 1.14.17.1]. The Cu-containing glycoprotein enzyme has been isolated by two procedures. The first is an elaborate method requiring extraction, two  $(\text{NH}_4)_2\text{SO}_4$  fractionations, calcium phosphate gel filtration, EtOH fractionation, DEAE-cellulose chromatography followed by two Sephadex-G200 gel filtrations giving enzyme with a specific activity of 65 Units/mg. [Friedman & Kaufman *J Biol Chem* **240** 4763 1965, Rush et al. *Biochem Biophys Res Commun* **61** 38 1974.] The second procedure is much gentler and provides good quality enzyme. Sedimented chromaffin vesicles are lysed in 10 volumes of 5mM K-phosphate buffer pH 6.5 using a loosely fitting Teflon-glass homogeniser. The mixture is centrifuged at 40,000xg/0.5 hours, and the supernatant is diluted with an equal volume of 100mM phosphate buffer (pH 6.5) containing 0.4M NaCl. This lysate is applied to a concanavalin A-Sepharose column (4 x 0.7cm) which had been equilibrated with 50 mM of phosphate buffer (pH 6.5 + 0.2M NaCl) with a flow rate of  $\sim 0.3$  ml/minute. The column is washed thoroughly with the buffer until  $\text{OD}_{280\text{nm}}$  is 0.005. The enzyme is then eluted with the same buffer containing 10%  $\alpha$ -methyl-D-mannoside (flow rate 0.1 ml/minute), and the enzyme is collected in 20-column volumes. The pooled eluate is concentrated by ultrafiltration in an Amicon Diaflo stirrer cell using an XM100A membrane. The concentrated enzyme is dialysed against 50mM phosphate buffer (pH 6.5) containing 0.1% NaCl. The enzyme gives one band (+ two very weak bands) on disc gel electrophoresis indicating better than 93% purity (67% fold purification) and has a specific activity of 5.4 Units/mg. [Rush et al. *Biochem Biophys Res Commun* **57** 1301 1974, Stewart & Klinman *Ann Rev Biochem* **57** 551 1988.]

**Exonucleases.** Like the endonucleases they are restriction enzymes which act at the 3' or 5' ends of linear DNA by hydrolysing off the nucleotides. Although they are highly specific for hydrolysing nucleotides at the 3' or 5' ends of linear DNA, the number of nucleotides cleaved is time dependent and usually has to be estimated from the time allocated for cleavage. Commercially available exonucleases are used without further purification.

**Ferritin (from human placenta)** [9007-73-2]  $M_r \sim 445,000$  (Fe free protein). The purification of this major iron-binding protein is achieved by homogenisation in water and precipitation with ammonium sulfate, repeating the cycle of ultracentrifugation, and molecular sieve chromatography through a Sephadex 4B column. Isoelectric focusing reveals a broad spectrum of impurities which can be separated by ion-exchange chromatography on Sephadex A-25 and stepwise elution. [Konijn et al. *Anal Biochem* **144** 423 1985.]

**Fibrinogen (from human plasma)** [9001-32-5]  $M_r$  341,000. This protein is made up of  $2\alpha_1, 2\beta$  and  $2\gamma$  subunits connected by disulfide bridges. A likely impurity is plasminogen. It is purified by glycine precipitation [Mosesson & Sherry *Biochemistry* **5** 2829 1966] to obtain fractions 1-2, then further purified [Blombäck & Blombäck *Arkiv Kemi* **10** 415 1956] and contaminating plasminogen is removed by passage through a lysine-Sepharose column. Such preparations are at least 95% clottable as determined by Mosesson and Sherry's method (above ref.) in which the  $\text{OD}_{280}$  is measured before and after clotting with 5 Units/ml of thrombin ( $>3000\text{U/mg}$ ). All fibrinogen preparations are treated with calf intestinal alkaline phosphatase to convert any fibrinogen peptide-AP to fibrinogen peptide-A by removing serine-bound phosphate. Solutions are then lyophilised and stored at  $-20^\circ$ . [Higgins & Shafer *J Biol Chem* **256** 12013 1981.] It is sparingly soluble in  $\text{H}_2\text{O}$ . Aqueous solutions are viscous with isoelectric point at pH 5.5. It is readily denatured by heating above  $56^\circ$  or by chemical agents, e.g. salicylaldehyde, naphthoquinone sulfonates, ninhydrin or alloxan. [Edsall et al. *J Am Chem Soc* **69** 2731 1947, Purification: Cama et al. *Naturwissenschaften* **48** 574 1961, Lorand & Middlebrook

*Science* **118** 515 1953, cf. Fuller in *Methods Enzymol* **163** 474 1988.]

For plasminogen-deficient fibrinogen from blood plasma, the anticoagulated blood is centrifuged and the plasma is frozen and washed with saline solution. It is treated with charcoal, freeze-thawed and dialysed *versus* Tris/NaCl buffer. [Maxwell & Nikel *Biochemical Preparations* **12** 16 1968.]

**Fibronectin (from human plasma)** [86088-83-7]  $M_r \sim 220,000$ . This glycoprotein contains 5-12% of carbohydrate. It has been purified by glycine fractionation and DEAE-cellulose chromatography. This material is dissolved in 0.25M Tris-phosphate buffer pH 7.0, diluted to 20% and glycine added gradually till 2.1M when the temperature falls to below 15°. The precipitate contains mainly fibrinogen. The supernatant is discarded, and the precipitate is treated with an equal volume of H<sub>2</sub>O, cooled (to 0°) and precipitated by adding EtOH to 16% (v/v) at -4°. The precipitate contains some CI (Cold Insoluble) globulin, fibronectin and small quantities of other proteins. To remove these the precipitate is dissolved in 0.25M Tris-phosphate buffer (pH 7.0) *ca* 0.5% and purified by DEAE-cellulose chromatography after diluting the buffer to 0.05M buffer. [Morrison et al. *J Am Chem Soc* **70** 3103 1948, Mosesson & Umfleet *J Biol Chem* **245** 5728 1970, Mosesson & Amrani *Blood* **56** 145 1980, Akiyama & Yamada *Adv Enzymol* **59** 51 1987.]

**Follicle Stimulating Hormone (FSH, follitropin)** [9002-68-0]  $M_r \sim 36,000$ . FSH is purified by Sephadex G100 gel filtration followed by carboxymethyl-cellulose with NH<sub>4</sub>OAc pH 5.5. The latter separates luteinising hormone from FSH. Its solubility in H<sub>2</sub>O is 0.5%. It has an isoelectric point of 4.5. A solution of 1mg in saline (100ml) can be kept at 60° for 0.5 hour. Activity is retained in a solution at pH 7-8 for 0.5 hour at 75°. The activity of a 50% aqueous EtOH solution is destroyed at 60° in 15 minutes. [Bloomfield et al. *Biochim Biophys Acta* **533** 371 1978, Hartree *Biochem J* **100** 754 1966, Pierce & Parsons *Ann Rev Biochem* **50** 465 1981.]

**β-Galactosidase (from bovine testes)** [9031-11-2]  $M_r$  510,000 [EC 3.2.1.23]. It is purified 600-fold by (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> precipitation, acetone fractionation and affinity chromatography on agarose substituted with terminal thio-β-galactopyranosyl residues. [Distlern & Jourdan *J Biol Chem* **248** 6772 1973.]

**Glucose oxidase (from *Aspergillus niger*)** [9001-37-0]  $M_r$  186,000, [EC 1.1.3.4]. The oxidase is purified by dialysis against deionized water at 6° for 48 hours and by molecular exclusion chromatography with Sephadex G-25 at room temperature. [Holt & Cotton *J Am Chem Soc* **109** 1841 1987.]

**Glucose-6-phosphate dehydrogenase** [9001-40-5]  $M_r$  128,000 (from Baker's yeast), 63,300 (from rat mammary gland) [EC 1.1.1.49]. The enzyme is useful for measuring pyridine nucleotides in enzyme recycling. The enzyme from Baker's yeast has been purified by (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation, Me<sub>2</sub>CO precipitation, a second (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation, concentration by DEAE-SF chromatography, a third (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation and recrystallisation. Crystallisation is induced by addition of its coenzyme NADP, which in its presence causes rapid separation of crystals at (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> concentration much below that required to precipitate the amorphous enzyme. To recrystallise, the crystals are dissolved in 0.01M NADP (pH 7.3) with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> at 0.55 saturation, and the crystals appear within 10 to 60 minutes. After standing for 2-3 days (at 4°) the (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is increased to 0.60 of saturation, and more than 80% of the activity in the original crystals is recovered in the fresh crystals. [Noltmann et al. *J Biol Chem* **236** 1255 1961]. Large amounts can be obtained from rat livers. The livers are extracted with 0.025M phosphate buffer (pH 7.5) and precipitated with 3M (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (70% of activity). The precipitate is dissolved in 3 volumes of 0.025M phosphate (pH 7.5), dialysed against this buffer + 0.2mM EDTA at 4° for 5 hours, then diluted to 1% protein and the nucleic acids are precipitated by addition of 0.4 volumes of 1% protamine sulfate. (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is added to a concentration of 2M (pH adjusted to 7.0 with NH<sub>3</sub>), the precipitate is discarded and the supernatant is adjusted to 2.8M (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, dialysed, and the protein is adjusted to 1% and treated with Ca<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub> gel. The gel is added in three steps (1.5ml of 0.4% gel/ml per step), and the gel is removed by centrifugation after each addition. The third gel adsorbed 50% of the activity. The gel is eluted with 0.2M phosphate buffer (pH 7.4, 40ml/g of gel; 60% recovery). The extract is precipitated in 3 volumes of (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (adjusted to 4M) to give enzyme with an activity of 30μmoles/mg of protein per hour. [Lowry et al. *J Biol Chem* **236** 2746 1961.] The  $K_m$  values for the yeast enzyme are 20μM for G-6P and

2 $\mu$ M for NADP (Tris pH 8.0, 10<sup>-2</sup> M MgCl<sub>2</sub>, 38°) [Noltmann & Kuby *The Enzymes* VII 223 1963].

**Glutathione S-transferase (human liver)** [50812-37-8]  $M_r$  25,000 [EC 2.5.1.18]. It is purified by affinity chromatography using a column prepared by coupling glutathione to epoxy-saturated Sepharose. After washing contaminating proteins, the pure transferase is eluted with buffer containing reduced glutathione. The solution is then concentrated by ultrafiltration, dialysed against phosphate buffer at pH ~7 and stored in the presence of dithiothreitol (2mM) in aliquots at < -20°. [Simons & Vander Jag *Anal Biochem* 52 334 1977.]

**Glyceraldehyde-3-phosphate dehydrogenase** [9001-50-7]  $M_r$  144,000 [EC 1.2.1.12]. Purify the dehydrogenase from rabbit muscle by extraction with 0.03N KOH and precipitate it with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (0.52 of saturation). The clear supernatant is adjusted to pH 7.5, and NH<sub>3</sub> is added dropwise to pH 8.2-8.4. Crystals appear sometimes even without seeding. The crystals are dissolved in H<sub>2</sub>O, filtered to remove suspended material and 2 volumes of saturated (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> at pH 8.2-8.4 is added. After 1 hour the crystals appear. Recrystallise it in the same way. [Cori et al. *J Biol Chem* 173 605 1948, Furfine & Velick *J Biol Chem* 240 844 1965, *The Enzymes* 7 243 1963, Lui & Huskey *Biochemistry* 31 6998 1992.] The Km values are: NADH (3.3 $\mu$ M) and 1,3-diphosphoglycerate (8x10<sup>-7</sup>M) in pH 7.4 imidazole buffer at 26°, NAD (13 $\mu$ M), glyceraldehyde-3-P (90 $\mu$ M), P<sub>i</sub> (2.9x10<sup>-4</sup>M), and arsenate (69 $\mu$ M) in 8.6 M NaHCO<sub>3</sub> buffer at 26°C. [Orsi & Cleland *Biochemistry* 11 102 1972.]

**Glycerol kinase (from *Candida mycoderma*, *E coli*, rat or pigeon liver glycerokinase)** [9030-66-4]  $M_r$  251,000 [EC 2.7.1.30]. Commercial enzyme has been dialysed against 2mM Hepes, 5mM dithiothreitol and 0.3mM EDTA, followed by several changes of 20mM Hepes and 5mM dithiothreitol prior to storage under N<sub>2</sub> at -20°. [Knight & Cleland *Biochemistry* 28 5728 1989.] The enzyme from pigeon liver is purified by acid-precipitation (acetate buffer at pH 5.1), (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation, heat treatment (60°/ 1 hour), calcium phosphate gel filtration, a second (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation, dialysis, elution of inert proteins and crystallisation. This is done by repeatedly extracting the precipitate from the last step with 0.05M sodium pyrophosphate (pH 7.5) containing 1mM EDTA and 0.2M (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is added. Careful addition of solid (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> to this solution leads to crystallisation of the enzyme. Recrystallisation is repeated. The enzyme is activated by Mg<sup>2+</sup> and Mn<sup>2+</sup> ions and is most stable in solutions in the pH 4.5-5.5 range. The stability is greatly increased in the presence of glycerol. It has Km for glycerol of 60 $\mu$ M and for ATP 9 $\mu$ M in glycine buffer pH 9.8 and 25°. [Kennedy *Methods Enzymol* 5 476 1962.]

**L-Glycerol-3-phosphate dehydrogenase (GDH, from rabbit muscle)** [9075-65-4]  $M_r$  78,000 [EC 1.1.1.8]. The dehydrogenase is recrystallised by adding (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> to 0.45 saturation at pH 5.5 at 4°C, and the small amount of precipitate is removed, then a saturated solution of (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> is added dropwise from time to time over several days in the cold room. The crystals are collected and recrystallised until they have maximum activity. The enzyme is stable in half-saturated (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> for several weeks at 4°. The equilibrium [dihydroxyacetone][NADH][H<sup>+</sup>]/[G-3-P][NAD] is 1.0 x 10<sup>-12</sup>M in Tris buffer at 25°. It uses NAD ten times more efficiently than NADP. The Km for G-3-P is 1.1 x 10<sup>-4</sup>M, for NAD it is 3.8 x 10<sup>-4</sup>M and for dihydroxyacetone it is 4.6 x 10<sup>-4</sup>M in phosphate buffer pH 7.0 and at 23.3°. Dihydroxyacetone phosphate and fructose-1,6-diphosphate are inhibitors. [Branowski *J Biol Chem* 180 515 1949, *The Enzymes* 7 85 1963, Young & Pace *Arch Biochem Biophys* 75 125 1958, Walsh & Sallach *Biochemistry* 4 1076 1965.]

**Glycogen synthase (from bovine heart)** [9014-56-6]  $M_r$  60,000 [EC 2.4.1.11]. Purify the synthase by precipitation in the presence of added glycogen by polyethylene glycol, chromatography on DEAE-Sephacel and high-speed centrifugation through a sucrose-containing buffer. [Dickey-Dunkirk & Kollilea *Anal Biochem* 146 199 1985.]

**Haemoglobin A (from normal human blood)** [9008-02-0]  $M_r$  ~64,500, amorphous. Purify it from blood using CM-32 cellulose column chromatography. [Matsukawa et al. *J Am Chem Soc* 107 1108 1985.] For the purification of the  $\alpha$  and  $\beta$  chains see Hill et al. *Biochemical Preparations* 10 55 1963.

**Histones (from S4A mouse lymphoma).** The purification of histones uses a macroprocess column, heptafluorobutyric acid as solubilising and ion-pairing agent and an acetonitrile gradient. [McCroskey et al. *Anal Biochem* **163** 427 1987.]

**Hyaluronidase** [9001-54-1, 37326-33-3]  $M_r$  **43,000 (bovine testes), 89,000 (bacterial)** [EC 3.2.1.35]. Hyaluronidase is purified by chromatography on DEAE-cellulose prior to use. [Distler & Jourdain *J Biol Chem* **248** 6772 1973.]

**D-Hydantoinase** [dihydropyrimidinase, also called 5,6-dihydropyrimidine amidohydrolase, from microorganisms e.g. *Pseudomonas*, *Bacillus*, *Agrobacterium* as well as from mammalian and human tissues] [9030-74-4]  $M$  **51,720 (monomer from amino acid sequence, usually dimer or tetramer)** [EC 3.5.2.2],  $pI$  **~6.5**. This cytosolic enzyme hydrolyses dihydropyrimidines and hydantoins to N-carbamoylamino acids, and with the appropriate substrates are useful for preparing D-amino acids. The enzyme from a recently isolated species of *Agrobacterium* was purified to homogeneity and found to possess hydantoinase activity that was free from dihydropyrimidinase activity. It had an estimated subunit molecular weight of ~66,500 and a theoretical molecular weight of 265,000. The preferred substrates were 5-mono-substituted hydantoins with aromatic groups as shown from the  $K_m$  values. 5,5-Dimethylhydantoin and thio analogues of 5-*p*-hydroxyphenylhydantoins were competitive inhibitors. [Runser & Meyer *Eur J Biochem* **213** 131 1993.]

A commercially available hydantoinase preparation from *Vigna angularis* from aduki bean with a minimum activity of 300U/g is commercially available. These enzymes are generally inhibited by N-carbamoylamino acids (reaction products) 8-hydroxyquinoline, EDTA,  $Sn^{2+}$  and  $Zn^{2+}$ , but are activated by uracil, 2-thiouracil,  $Co^{2+}$ ,  $Fe^{2+}$ ,  $Mg^{2+}$ ,  $Mn^{2+}$  and  $Ni^{2+}$ , with turnovers of ~27,000min<sup>-1</sup> for hydantoin. Their pH-range is 6—9.5, and the temperature range is ~4-60°. D-Hydantoinase genes have been cloned and expressed in *E coli* and the enzymes from several sources have been crystallised.

A unit of enzyme activity is defined as the amount that catalyses the formation of 1mmole of N-carbamoylglycine from hydantoin per minute at pH 9.0 and 40°. [Morin *Enz Microbiol Technol* **15** 208 1993, *Springer Handbook of Enzymes* D. Schonburg & I. Schonburg Eds (A. Chang co-Ed) Springer-Verlag, Berlin, Heidelberg.]

**3-Hydroxy butyrate dehydrogenase (from *Rhodopseudomonas spheroides*)** [9028-38-0]  $M_r$  **~85,000** [EC 1.1.1.30], **amorphous**. Purify the dehydrogenase by two sequential chromatography steps on two triazine dye-Sepharose matrices. [Scavan et al. *Biochem J* **203** 699 1982.]

**Interferons [ $\alpha$ IFN,  $\beta$ IFN and  $\gamma$ IFN].** Interferons are a family of glycosylated proteins and are cytokines which are produced a few hours after cells have been infected with a virus. Interferons protect cells from viral infections and have antiviral activities at very low concentrations (~3 x 10<sup>-4</sup> M; less than 50 molecules are apparently sufficient to protect a single cell). Double-stranded RNAs are very efficient inducers of IFNs. There are three main types of IFNs. The  $\alpha$ IFNs are synthesised in lymphocytes, and the  $\beta$ IFNs are formed in infected fibroblasts. The  $\alpha$  and  $\beta$  families are fairly similar, consisting of *ca* 166 to 169 amino acids. Although  $\gamma$ IFNs are also small glycosylated proteins (*ca* 146 amino acids), they are different because they are not synthesised after viral infections but are produced by lymphocytes when stimulated by **mitogens** (agents that induced cell division).

Several of these IFNs of mouse and human lymphocytes and fibroblasts are available commercially and have been best prepared in quantity by recombinant DNA procedures because they are produced in very small amounts by the cells. The commercial materials do not generally require further purification for their intended purposes. [Pestkas, Interferons and Interferon standards and general abbreviations, *Methods Enzymol*, Wiley & Sons, **119** 1986, ISBN 012182019X, Lengyel, Biochemistry of interferons and their actions, *Ann Rev Biochem* **51** 251-282 1982, De Maeyer & De Maeyer-Guignard, Interferons in *The Cytokine Handbook*, 3rd Edn, Thomson et al. Eds, pp. 491-516 1998 Academic Press, San Diego, ISBN 0126896623.]

**Interleukin (from human source).** Purify these using lyophilisation and desalting on a Bio-Rad P-6DC desalting gel, then two steps of HPLC, first with hydroxylapatite, followed by a TSK-125 size exclusion column. [Kock & Luger *J Chromatogr* **296** 293 1984.]

**Interleukin-2 (recombinant human)** [94218-72-1]  $M_r$  ~15,000, amorphous. Purify it by reverse phase HPLC. [Weir & Sparks *Biochem J* **245** 85 1987, Robb et al. *Proc Natl Acad Sci USA* **81** 6486 1984.]

**Interleukins (IL-1, IL-2 —IL18).** Interleukins are cytokines which cause a variety of effects including stimulation of cell growth and proliferation of specific cells, e.g. stem cells, mast cells, activated T cells, colony stimulating factors etc., as well as stimulating other ILs, prostaglandins release etc. They are small glycosylated proteins (ca 15 kD, 130-180 amino acids produced from longer precursors) and are sometimes referred to by other abbreviations, e.g. IL-2 as TCGF (T cell growth factor), IL-3 as multi-CSF (multilineage colony stimulating factor, also as BPA, HCSF, MCSF and PSF). They are produced in very small amounts and are commercially made by recombinant DNA techniques in bacteria or Sf21 insect cells. Interleukins for human (h-IL), mouse (m-IL) and rat (r-IL) are available, and up to IL-18 are available commercially in such purity that they can be used directly without further refinement, particularly those that have been obtained by recombinant DNA procedures which are specific. As well as the interleukins, a variety of antibodies for specific IL reactions are available for research or IL identification. [Symons et al. *Lymphokines and Interferons, A Practical Approach*, Clemens et al. Eds, p.272 1987, IRL Press, Oxford, ISBN 1852210354, 1852210362, Thomson et al. Eds, *The Cytokine Handbook*, 3rd Edn, 1998, Academic Press, San Diego, ISBN 0126896623.]

**Lactate dehydrogenase (from dogfish, Beef muscle)** [9001-60-9]  $M_r$  140,000 [EC 1.1.1.27]. A forty-fold purification of the dehydrogenase is effected by affinity chromatography using Sepharose 4B coupled to 8-(6-aminohexyl)amino-5'-AMP or -NAD<sup>+</sup>. [Lees et al. *Arch Biochem Biophys* **163** 561 1974, Pesce et al. *J Biol Chem* **239** 1753 1964.]

**Lactoferrin (from human whey)** [55599-62-7, Fe Salt]  $M_r$  ~90,000. This iron-binding protein is purified by direct adsorption on cellulose phosphate by batch extraction, then eluted by a stepped salt and pH gradient. The Fe bound protein forms red crystals with  $\lambda_{max}$  at 465nm (pH 8.2). [Foley & Bates *Anal Biochem* **162** 296 1987.]

**Lectins (proteins and/or glycoproteins of non-immune origin that agglutinate cells, e.g. from seeds of *Robinia pseudoacacia*),**  $M_r$  ~100,000. Lectins are purified by precipitation with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> and dialysed, then chromatographed on DE-52 DEAE-cellulose anion-exchanger, hydroxylapatite and Sephacryl S-200. [Wantyghem et al. *Biochem J* **237** 483 1986.]

Lectins are a group of proteins that are classed as sugar-binding proteins or glycoproteins of non-immune origin and which agglutinate cells and/or precipitate glyco-conjugates. They are present in plants (seeds, roots, leaves or bark) and some invertebrates (snails, clams, crabs) and have  $M_r$  values of 10,000-400,000. They may contain Mn<sup>2+</sup> and/or Ca<sup>2+</sup>. Mono- or oligo- saccharides of appropriate specificity inhibit lectins. Some lectins are specific to human blood groups and induce mitosis in lymphocytes. [Goldstein *Nature* **286** 66 1980.]

**Lipoprotein lipase (from bovine skimmed milk)** [9004-02-8]  $M_r$  ~34,000 and 63,000 (SDSPAGE), 96,900 (sedimentation and diffusion), 100,000-120,000 (gel filtration) [EC 3.1.1.34]. Purify the lipase by affinity chromatography on heparin-Sepharose;  $K_i$  0.026mM for very low density lipoprotein. It is inhibited by 2-mercaptoethanol, Cys, Ca, Hg, Mg and Mn ions. Protamine sulfate, 1mg of bovine serum albumin/ml or in 50% glycerol at -70°, stabilises the lipase for several days. 60% loss of activity occurs at 0°/1hour in the presence of 1% of bovine serum albumin. [Shirai et al. *Biochim Biophys Acta* **665** 504 1981.]

**Lipoproteins (from human plasma).** Individual human plasma lipid peaks are removed from plasma by ultracentrifugation; then they are separated and purified by agarose-column chromatography. Fractions are characterised immunologically, chemically, electrophoretically and by electron microscopy. [Rudel et al. *Biochem J* **13** 89 1974.]

**Lipoteichoic acids (from gram-positive bacteria)** [56411-57-5]. The acids are extracted by hot phenol/water from disrupted cells. Nucleic acids are also extracted and are removed by treatment with nucleases. Nucleic resistant acids, proteins, polysaccharides and teichoic acids are separated from lipoteichoic acids by anion-exchange chromatography on DEAE-Sephacel or by hydrophobic interaction on octyl-Sepharose [Fischer et al. *Eur J Biochem* **133** 523 1983].

**Lysozyme (Muramidase, N-acetylmuramyl hydrolase, N-acetylmuramide, peptidoglycan N-acetylmuramoyl hydrolase, Globulin G<sub>1</sub>)** [human 174883-18-2, from human neutrophils 9001-63-2, from human milk 12671-19-1, from chicken egg white 126050-88-3, chloride from chicken egg white 9066-59-5] **M 14,400 ±100, E<sub>280nm</sub> 2.65 (c 1mg/ml) Isoelectric Point (IP) 10.5-11.0 [EC 3.2.1.17]**. Lysozymes from human and bird sources are 129 amino acid enzymes that contain four disulfide bonds. Hen lysozyme was isolated in quantity from chicken egg white, crystallised, and its X-ray crystalline structure was determined [Phillips *Proc Natl Acad Sci USA* **57** 484 1967]. Lysozyme occurs in many tissues of invertebrate and vertebrate animals. It is found in milk, blood serum and various secretions (saliva, nasal mucus and tears). It also occurs in some moulds and in the latex of certain plants. Lysozyme from human milk was studied in some detail and is very similar to the hen enzyme [Jolles & Jolles *Hely Chim Acta* **52** 2671 1969 and **54** 2668 1971]. Lysozyme is a glycosidase which dissolves various bacterial cell walls [particularly Gram-positive bacteria (which have surface lipoproteins) to give spheroblasts, and various Gram-negative bacteria in the presence of EDTA in hypotonic solutions, or non-ionic detergents]. Bacterial cell walls contain 1,4-β-N-acetylglucosaminyl oligosaccharides which are cleaved by the enzyme at the glycosidic C—O bond between the 4th and 5th sugar residues from the non-reducing end of the chain. The mechanism and kinetics of this hydrolysis have been studied extensively [cf Fersht *Enzyme Structure and Mechanism, 2nd edn*, W.H. Freeman & Co, Reading 1985, ISBN 0716716151]. Lysozyme was an extremely useful antibacterial in the pre-antibiotic era. It is a basic protein with 20-22 basic residues and only 3-4 acidic groups (see their isoelectric points); it forms soluble salts and is stable up to 55°. It is purified from egg white by chromatography through Amberlite IRC50 at pH 7.18 in 0.2M phosphate buffer followed by recrystallisation at pH 9.5 by adding NaCl to a concentration of 5%. Alternatively, ~1L of homogenised egg white in the absence of air (from 3 dozen eggs) is added to a 10% suspension of Bentonite [1302-78-9] (150ml, a native hydrated aluminium silicate) in 1% aqueous KCl, and stirred vigorously enough to avoid excessive foaming for ~5 minutes to give a smooth suspension. Separate the clay by centrifugation, wash it twice with 0.5M phosphate buffer (300ml each, pH 7.5) and three times with 5% aqueous pyridine (300ml each) while decanting from the clay and discarding inactive supernatants. From the combined washed clay, lysozyme is extracted out twice with 300ml of 5% aqueous pyridine (adjusted to pH5 with H<sub>2</sub>SO<sub>4</sub> using a glass electrode). The combined extracts are dialysed against running tap H<sub>2</sub>O (at 12-15°) until free from pyridine odour and dialysed further for 24 hours. Essentially pure amorphous lysozyme is obtained by lyophilising the dialysate below 25°. *Crystalline isoelectric lysozyme* is obtained from the amorphous powder by adding 0.5g of NaCl to a 5% aqueous solution of the enzyme (10ml), adjusting the pH to 9.5-10 (glass electrode) with aqueous NaOH and storing at 4°, whereby it crystallises out. *Crystalline lysozyme carbonate* is obtained from the amorphous powder by adding 0.5g of NaHCO<sub>3</sub> to a 5% aqueous solution of the enzyme (10ml), giving a final pH of 8.0-8.5 (glass electrode) with aqueous NaOH and allowing to stand at room temperature, whereby the carbonate crystallises out. *Crystalline lysozyme chloride* is obtained from a 5% solution of amorphous powder by adjusting the pH to 4.5 with hydrochloric acid (glass electrode) followed by solid NaCl to 5%, and storing at 4°, whereby the chloride crystallises out in 4 to 5 days. [Alderton & Fevold *J Biol Chem* **164** 1 1946, Fevold & Alderton *Biochemical Preparations* **I** 67 1949.] Other salts include the *L-ascorbate* [119189-24-1], the *lactate* [72497-48-4], the *phosphate* [72497-47-3] and the *2,3-dihydroxypropyl dihydrogen phosphate* [119189-23-0]. Dialysis and recrystallisation are simple and yield enzyme of high purity. Several forms of crystals are obtained depending on the pH of the crystallising solution. The activity of lysozyme is measured by the rate of decrease of turbidity (at 570nm) as hydrolysis of acetone-dried cell walls of the Gram-positive bacterium *Micrococcus lysodeikitus* (as substrate) occurs on addition of the enzyme [Hirs *Methods Enzymol* **I** 124 1968]. At high concentrations or in the presence of albumin, lysozyme can be lyophilised or desalted without loss of activity. Freezing and thawing does not inactivate it, and Na<sup>+</sup>, Mg<sup>2+</sup>, Ag<sup>+</sup>, Ca<sup>2+</sup>, and Cu<sup>2+</sup> exert an activating effect. It can be stored for several weeks at -20° without inactivation, and in the presence of 0.1% of bovine serum albumin it can be stored for months at -20° without inactivation. [Springer *Handbook of Enzymes* 2nd edn. Schonburg & Schonburg eds (A. Chang co-ed) Springer-Verlag Berlin, Heidelberg, **Volume 12** 2003 ISBN 3-540-00519-6.] Agarose-bound lysozyme from hen egg (5,000-10,000 Units/g of solid), a lysozyme Biotin-caproyl solution (20,000 U/mg) and a carboxymethylated-maleylated lysozyme (Lysozyme-RCM, reduced form), and lysozyme of up to 100,000 U/mg protein are commercially available.

A *T<sub>4</sub> bacteriophage lysozyme* (from phage grown on *E.coli* B<sup>1</sup>) [12585-29-4] was extracted and freed from

DNA, cell debris, intact cells and acidic proteins by precipitation with Rivanol (6,9-diamino-2-ethoxyacridine *dl*-lactate). The filtrate was purified by concentration through an Amberlite IRC50 column which was thoroughly washed with 0.1M phosphate buffer pH 6.5 containing  $10^{-3}$ M  $MgSO_4$  and eluted with the phosphate buffer at pH 5.8 containing a 0 to 0.5M NaCl gradient. This was repeated twice, followed by gel filtration through Sephadex G-75 in pH 5.8 phosphate buffer and eluted with 0.5M NaCl. The eluate was dialysed against  $H_2O$  and lyophilised to give a 1500-fold purification with 40% recovery. Like the hen lysozyme it is a basic protein but with 164 amino acid residues, and it is unstable  $>40^\circ$ . [Tsugita et al. *J Biol Chem* **243** 391 1968, Inouye et al. *J Biol Chem* **245** 3439, 3479 1970, Matthews *Biochim Biophys Acta* **405** 442 1975.]

**Metallothionein (from rabbit liver)** [9038-94-2]. Purify it by precipitation to give Zn- and Cd-containing protein fractions and running it on a Sephadex G-75 column, then isoelectric focusing to give two protein peaks [Nordberg et al. *Biochem J* **126** 491 1972, Comeau et al. *Prep Biochem* **22** 151 1992].

**Myoglobin (from sperm whale muscle)** [9047-17-0]  $M_r \sim 17,000$ . Myoglobin is purified by CM-cellulose chromatography and Sephadex G-50 followed by chromatography on Amberlite IRC-50 Type III or BioRex 70 ( $<400$ mesh). The crystalline product as a paste in saturated  $(NH_4)_2SO_4$  at pH 6.5-7.0 may be stored at  $4^\circ$  for at least 4 years unchanged, but must not be kept in a freezer. [Anres & Atassi *Biochemistry* **12** 942 1980, Edmundson *Biochemical Preparations* **12** 41 1968.]

**5'-Nucleotidase (from Electric ray, *Torpedo sp*)** [9027-73-0] [EC 3.1.3.5], **amorphous**. Purify it by dissolving it in Triton X-100 and deoxycholate, and by affinity chromatography on concanavalin A-Sepharose and AMP-Sepharose [Grondal & Zimmerman *Biochem J* **245** 805 1987].

**Orosomucoid (glycoprotein  $\alpha_1$  acid, from human plasma)** [66455-27-4]  $M_r$  **42000-44000**, **amorphous**. Purify the glycoprotein  $\alpha_1$  acid by passage through a carboxymethyl cellulose column, then through a Sephadex G-25 column. [Aronson et al. *J Biol Chem* **243** 4564 1968.]

**Papain** [9001-73-4]  $M_r \sim 21,000$  [EC 3.4.22.2], **amorphous**. A suspension of 50g of papain (freshly ground in a mortar) in 200ml of cold water is stirred at  $4^\circ$  for 4 hours, then filtered through a Whatman No 1 filter paper. The clear yellow filtrate is cooled in an ice-bath while a rapid stream of  $H_2S$  is passed through it for 3 hours, and the suspension is centrifuged at 2000rpm for 20 minutes. Sufficient cold MeOH is added slowly with stirring to the supernatant to give a final MeOH concentration of 70 vol%. The precipitate, collected by centrifugation for 20 minutes at 2000rpm, is then dissolved in 200ml of cold water, the solution is saturated with  $H_2S$ , centrifuged, and the enzyme is again precipitated with MeOH. The process is repeated four times. [Bennett & Niemann *J Am Chem Soc* **72** 1798 1950.] Papain has also been purified by affinity chromatography on a column of GlyGlyTyrArg-agarose [Stewart et al. *J Am Chem Soc* **109** 3480 1986].

**Pepsin** [9001-75-6]  $M_r$  **31,500(human), 6000(hog)** [EC 3.4.23.1]. Pepsin is re-chromatographed on a column of Amberlite CG-50 using a pH gradient prior to use. Crystallise it from EtOH. [Richmond et al. *Biochim Biophys Acta* **29** 453 1958, Huang & Tang, *J Biol Chem* **244** 1085 1969, **245** 2189 1970.]

**Pertussis toxin (from *Bordetella pertussis*)** [70323-44-3]  $M_r$  **117,000**. Purify the toxin by stepwise elution from 3 columns comprising Blue Sepharose, Phenyl Sepharose and hydroxylapatite, and the purity is checked by SDS-PAGE. [Svoboda et al. *Anal Biochem* **159** 402 1986, *Biochemistry* **21** 5516 1982, *Biochem J* **83** 295 1978.]

**Phosphatase alkaline (alkaline phosphatase)** [9001-78-9]  $M_r \sim 40,000$  (**bovine liver**),  $\sim 140,000$  (**bovine intestinal mucosa**), **80,000** (*E. coli*), [EC 3.1.3.1]. The *E. coli* supernatant in sucrose (20%, 33mM) in Tris-HCl pH 8.0 is purified through a DEAE-cellulose column and recrystallised. To the column eluates in 0.125M NaCl is added  $MgCl_2$  (to 0.01M) and brought to 50% saturation in  $(NH_4)_2SO_4$  by adding the solid (0.20g/ml). The mixture is centrifuged to remove bubbles and is adjusted to pH 8.0 (with 2N NaOH). Saturated  $(NH_4)_2SO_4$  at pH 8.0 is added dropwise until the solution becomes faintly turbid ( $\sim 61\%$  saturation). It is set aside at  $\sim 25^\circ$

for 1 hour (turbidity will increase). The mixture is placed in an ice bath for several minutes when turbidity

disappears and a clear solution is obtained. It is then placed in a large ice bath at 0° (~5L) and allowed to warm slowly to room temperature in a dark room whereby crystals are formed appearing as a silky sheen. The crystals are collected by centrifugation at 25° if necessary. The crystalline solutions are stable at ~25° for many months. They can be stored at 0°, but are unstable when frozen. Cysteine at 10<sup>-3</sup>M and thioglycolic acid at 10<sup>-4</sup>M are inhibitory. This is reversed on addition of Zn<sup>2+</sup> ions. Many organic phosphates are good substrates for this phosphatase. [Molamy & Horecker *Methods Enzymol* **9** 639 1966, Torriani et al. *Methods Enzymol* **12b** 212 1968, Engstrom *Biochim Biophys Acta* **92** 71 1964.]

Alkaline phosphatase from rat osteosarcoma has been purified by Me<sub>2</sub>CO precipitation and chromatography on DEAE-cellulose, Sephacryl S-200, and hydroxylapatite. [Nair et al. *Arch Biochem Biophys* **254** 18 1987.]

**Phosphoproteins (various).** These are purified by adsorbing onto an iminodiacetic acid substituted agarose column to which are bound ferric ions. This chelate complex acts as a selective immobilised metal affinity adsorbent for phosphoproteins. [Muszyfiska et al. *Biochemistry* **25** 6850 1986.]

**5'-Phosphoribosyl pyrophosphate synthetase (from human erythrocytes, or pigeon or chicken liver)** [9015-83-2]  $M_r$  60,000 [EC 2.7.6.1]. It is purified 5100-fold by elution from DEAE-cellulose, fractionation with (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>, filtration on Sepharose 4B and ultrafiltration. [Fox & Kelley *J Biol Chem* **246** 5739 1971, Flaks *Methods Enzymol* **6** 158 1963, Kornberg et al. *J Biol Chem* **15** 389 1955.]

**Pituitary Growth Factor (from human pituitary gland)** [336096-71-0]. It is purified by heparin and copper affinity chromatography, followed by chromatography on carboxymethyl cellulose (Whatman 52). [Rowe et al. *Biochemistry* **25** 6421 1986.]

**Plasmids.** These are circular lengths of *double-stranded* DNA which invade bacteria or other cells, e.g. insect cells, yeast cells, and have sequences which are necessary for their replication using enzymes and other ingredients, e.g. nucleotides, present in the cells. They contain engineered, or already have, genes which produce enzymes that provide the cells with specific antibiotic resistance and are thus useful for selecting bacteria containing specific plasmids. Plasmids have been extremely useful in molecular biology since they can be very easily identified (from their size or the sizes of the DNA fragments derived from their restriction enzyme digests) and can be readily engineered *in vitro* (outside the cells). Genes coding for specific enzymes or other functional proteins can be inserted into these plasmids which have DNA sequences that allow the expression of large quantities of bacterial or non-bacterial (e.g. human) proteins. They have also been engineered in such a way as to produce “fusion proteins” (in which the desired protein is fused with a specific “reporter, marker or carrier protein” which will facilitate the isolation of the desired protein (e.g. by binding strongly to a nickel support), and then the desired protein can be cleaved from the eluted fusion protein and obtained in very pure form. A large number of plasmids with a variety of sequences for specific purposes are commercially available in very pure form. They can be used to infect cells and can be isolated and purified from cell extracts in large amounts using a number of available procedures. These procedures generally involve lysis of the cells (e.g. with alkaline sodium dodecylsulfate, SDS), separation from nuclear DNA, precipitation of plasmid DNA from the cell debris, adsorbing it on columns which specifically bind DNA, and then eluting the DNA from the column (e.g. with specific Tris buffers as recommended by the suppliers) and precipitating it (e.g. with Tris buffer in 70% EtOH at -70°C). The purity is checked on agarose gels (containing ethidium bromide to visualise the DNA) by electrophoresis. A large number of plasmids are now commercially available (see Clontech GmbH, <http://www.clontech.com>, Invitrogen <http://www.invitrogen.com>, among other suppliers) and used as vectors for bacterial, mammalian, yeast and insect cells, and for baculovirus expression.

**Protamine kinase (from rainbow trout testes)** [37278-10-7] [EC 2.7.1.70]. Partial purification of the kinase is achieved by chromatography through hydroxylapatite followed by biospecific chromatography on nucleotide coupled Sepharose 4B [the nucleotide is 8-(6-aminohexyl)amine coupled cyclic-AMP]. It requires Mg ions specifically but is inactivated by Ca, Mn, Cu, Zn ions and thiol compounds. [Jergil & Dixon *J Biol Chem* **245** 1425 1970, Jergil et al. *Biochem J* **139** 441 1974.]

**Protamine sulfate (from herring sperm)** [9007-31-2] [ $\alpha$ ]<sub>D</sub><sup>22</sup> -85.5° (saturated H<sub>2</sub>O), pK<sup>25</sup> 7.4-8.0.

Protamine sulfate is a strongly basic protein (white powder, see pK) used to precipitate nucleic acids from crude protein extracts. It dissolves to the extent of 1.25% in H<sub>2</sub>O. It is freely soluble in hot H<sub>2</sub>O but separates as an oil on cooling. It has been purified by chromatography on an IRA-400 ion-exchange resin in the SO<sub>4</sub><sup>2-</sup> form and is washed with dilute H<sub>2</sub>SO<sub>4</sub>. Eluates are freeze-dried under high vacuum below 20°. This method is used to convert proteamine and protamine hydrochloride to the sulfate. [UV: Rasmussen *Hoppe Seyler's Z Physiol Chem* **224** 97 1934, Ando & Sawada *J Biochem (Tokyo)* **49** 252 1961, Felix & Hashimoto *Hoppe Seyler's Z Physiol Chem* **330** 205 1963.]

**Protease nexin (from cultured human fibroblasts)** [148263-58-5]. It is purified by affinity binding of protease nexin to dextran sulfate-Sepharose. [Farrell et al. *Biochem J* **237** 707 1986.]

**Proteins.** These are usually naturally occurring (or deliberately synthesised in microorganisms, e.g. bacteria, insect cells, or animal tissues), and are composed of a large number of  $\alpha$ -S (L)amino acid residues (except for L-cysteine which has the R absolute configuration), selected from the 20 or so natural amino acids, in specific sequences and in which the  $\alpha$ -amino group forms an amide (peptide) bond with the  $\alpha$ -carboxyl group of the neighbouring amino acid. The number of residues is usually upwards of 100. Proteins with less than 100 amino acids are better referred to as **polypeptides**. Aqueous soluble proteins generally fold into ball-like structures mainly with hydrophilic residues on the outside of the "balls" and hydrophobic residues on the inside. Proteins can exist singly or can form dimers, trimers, tetramers etc., consisting of similar or different protein subunits. They are produced by cells for a large variety of functions, e.g. enzymology, reaction mediation as in regulation of DNA synthesis or chaperonins for aiding protein folding, formation of pores in membranes for transport of ions or organic molecules, or for intra or intercellular signalling etc. The purity of proteins can be checked in denaturing (SDS, sodium dodecylsulfate) or non-denaturing polyacrylamide gels using electrophoresis (PAGE), and staining appropriately (e.g. with Coomassie Blue, followed by silver staining for higher sensitivity). If the protein is partly impure, then it should be purified further according to the specific literature procedures for the individual protein (see specific proteins in the *Methods Enzymol*, Wiley series).

**Proteoglycans (from cultured human muscle cells).** These are separated by ion-exchange HPLC using a Bio-gel TSK-DEAE 5-PW analytical column. [Harper et al. *Anal Biochem* **159** 150 1986.]

**Prothrombin (Factor II, from equine blood plasma)** [9001-26-7] **M<sub>r</sub> 72,000**. Prothrombin is purified by two absorptions on a barium citrate adsorbent, followed by decomposition of the adsorbents with a weak carboxylic cation-exchanger (Amberlite IRF-97), isoelectric precipitation (pH 4.7-4.9) and further purification by chromatography on Sephadex G-200 or IRC-50. Finally it is recrystallised from a 1% solution adjusted to pH 6.0-7.0 and partial lyophilisation to ca 1/5 to 1/10th volume and set aside at 2-5° to crystallise. Occasionally seeding is required. [Miller *Biochemical Preparations* **13** 49 1971.]

**Prymnesin (toxic protein from phytoflagellate *Prymnesium parvum*)** [11025-94-8]. Prymnesin is purified by column chromatography, differential dissolution and precipitation in solvent mixtures and differential partition between diphasic mixtures. The product has at least 6 components as observed by TLC. [Ulitzur & Shilo *Biochim Biophys Acta* **301** 350 1970.]

**Pyruvate kinase isoenzymes (from *Salmonella typhimurium*)** [9001-59-6] **M<sub>r</sub> 64,000** [EC 2.7.1.40], **amorphous**. These are purified by (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> fractionation and gel filtration, ion-exchange and affinity chromatography. [Garcia-Olalla & Garrido-Pertierra *Biochem J* **241** 573 1987.]

**Renal dipeptidase (from porcine kidney cortex)** [9031-96-3] **M<sub>r</sub> 47,000** [EC 3.4.13.11]. The dipeptidase is purified 700-fold by homogenising the tissue, extracting with Triton X-100, elimination of insoluble material, and ion-exchange, size exclusion and affinity chromatography. [Hitchcock et al. *Anal Biochem* **163** 219 1987.]

**Restriction enzymes (endonucleases).** These are enzymes which cleave double-stranded DNA (linear or circular) at specific nucleotide sequences within the DNA strands which are then used for cloning (by ligating bits of DNA sequences together) or for identifying particular DNA materials, e.g. plasmids, genes etc. A very large number of restriction enzymes are now available commercially and are extensively used in molecular biology. They are highly specific for particular nucleotide arrangements and are sensitive to the reaction conditions, e.g. composition of the medium, pH, salt concentration, temperature etc, which have to be strictly adhered to.

Endonucleases are essentially of two types. One type cleaves the DNA producing *blunt ends*, i.e. the nucleotides are paired at each end. The second type cleaves the DNA leaving *hanging ends*, i.e. the nucleotides at the ends of each strand (usually four nucleotides) of the double strands protrude ahead of the paired nucleotides. The latter enzymes particularly allow ligation of the cleaved sequence to occur in a desired direction.

The enzymes do not require further purification, and the reaction conditions are also provided by the suppliers from which the necessary reaction media can also be purchased (see commercial catalogues).

**Reverse transcriptase (from avian or murine RNA tumour viruses)** [9068-38-6]  $M_r \sim 170,000$  [EC 2.7.7.49]. This enzyme produces the complementary DNA from the RNA (as template). These are purified by solubilising the virus with non-ionic detergents. Lysed virions are adsorbed on DEAE-cellulose or DEAE-Sephadex columns, and the enzymes are eluted with a salt gradient, then chromatographed on a phosphocellulose column and fractions with enzyme activity are eluted in a salt gradient. They are also purified from other viral proteins by affinity chromatography on a pyran-Sepharose column. [Verna *Biochim Biophys Acta* **473** 1 1977, Smith *Methods Enzymol* **65** 560 1980; see commercial catalogues for other transcriptases.]

**Ribonuclease (from human plasma)** [9001-99-4]  $M_r \sim 13,700$  [EC 3.1.27.5], **amorphous**. Purify it by  $(\text{NH}_4)_2\text{SO}_4$  fractionation, followed by PC cellulose chromatography and affinity chromatography (using Sepharose 4B to which  $(\text{G})_n$  is covalently bonded). [Schmukler et al. *J Biol Chem* **250** 2206 1975.]

**Ribozymes.** These are ribonucleic acids which act like protein enzyme in catalysing the making and breaking of peptide bonds as well as catalysing reactions and cleavage of DNA and RNA molecules. The short RNAs are being intensively studied (see RNAi below). MicroRNAs (miRNA, see also below) are ubiquitous and are also genetically produced. They are involved in numerous reactions from splicing RNA, e.g. mRNA (messenger RNA) to controlling transcription of DNA to RNA, and translation of RNA to protein. Each miRNA is capable of being involved in a small number to hundreds of interactions with nucleic acids and with proteins.

**Ricin (toxin from Castor bean *Ricinus communis*)** [A chain 96638-28-7, B chain 96638-29-8]  $M_r \sim 30,000$ , **amorphous**. Crude ricin, obtained by aqueous extraction and  $(\text{NH}_4)_2\text{SO}_4$  precipitation, is chromatographed on a galactosyl-Sepharose column with sequential elution of pure ricin. The second peak is due to ricin agglutinin. [Simmons & Russell *Anal Biochem* **146** 206 1985.] It is an inhibitor of protein synthesis. **EXTREMELY DANGEROUS, USE EXTREME CARE [instructions accompany product].**

**RNA (ribonucleic acids).** Ribonucleic acids are like DNA except that the 2'-deoxy-D-ribose moiety is replaced by a D-ribose moiety and the fourth nucleotide thymidylic acid (T) is replaced by uridylic acid (U). RNA does not generally form complete duplex molecules like DNA, i.e. it is generally monomeric, except in certain viruses. The two main classes of RNA are **messenger-RNA** (mRNA) and **transfer-RNA** (tRNA). Pre-mRNAs are transcribed from the DNA genes, and non-coding segments (the introns) are spliced out to give the mRNAs which code for specific proteins. There are many different tRNAs, at least one of which is linked to a specific  $\alpha$ -amino acid that binds to the mRNA *via* the ribosome (a set of proteins) to the RNA triplets (three nucleotides) which code for the particular  $\alpha$ -amino acids. An enzyme then joins the  $\alpha$ -amino acids of two adjacent tRNA- $\alpha$ -amino acid ribosome complexes bound to the mRNA to form a peptide bond. Thus peptide bonds and consequently polypeptides and proteins coded by the DNAs, *via* the respective mRNA, are produced. Martin et al. [*Biochem J* **89** 327 1963] purified RNA by dissolving it (5g) in 90ml of 0.1mM EDTA, then homogenised this with 90ml of 90% (w/v) phenol in water using a Teflon pestle. The suspension was stirred vigorously for 1 hour at  $\sim 20^\circ$ , then centrifuged for 1 hour at  $0^\circ$  at 25,000rpm. The lower (phenol) layer was extracted four times with 0.1mM EDTA, and the aqueous layers were combined, then made 2% (w/v) with respect to AcOK and 70% (v/v) with respect to EtOH. After standing overnight at  $-20^\circ$ , the precipitate was centrifuged down, dissolved in

50ml of 0.1mM EDTA, made 0.3M in NaCl and kept for 3 days at 0°. The purified RNA was then centrifuged down at 10000x g for 30 minutes, dissolved in 100ml of 0.1mM EDTA, dialysed at 4° against water, and freeze-dried. It was stored at -20° in a desiccator. Michelson [*J Chem Soc* 1371 1959] dissolved 10g of RNA in water, added 2M ammonia to adjust the pH to 7, then dialysed in Visking tubing against five volumes of water for 24 hours. The process was repeated three times; then the material after dialysis was treated with 2M HCl and EtOH to precipitate the RNA which was collected, washed with EtOH, ether and dried [see commercial catalogues for further examples]. RNAs are now routinely sequenced.

**RNAi (interfering RNA).** RNAi technology has exploded during the past five years and has become invaluable for exploring gene function, interaction and control. It has considerable potential in therapeutics. This came about from the discovery that a double-stranded RNA (**dsRNA**) complementary to the coding sequence of a muscle protein “turned off” the gene instead of enhancing the effect. It seemed that the dsRNA activated the cellular machinery to degrade the target mRNA. This led to the development and the use of interfering RNA (RNAi) to “silence” specific genes. This silencing of genes comes in various different forms. It can be brought about by *short interfering RNA (siRNA)*, *micro RNA (miRNA)* and *short-hairpin RNA (shRNA)*. RNA can be notoriously unstable in biological systems, e.g. its half-life in serum is ~6 minutes, but recently methods have been devised (using nanoparticle delivery systems) where they are more stable and can be used to target specific cancer cells.

About 50% of the human genome is apparently transcribed into mRNA and translated into protein or used to produce small RNA fragments. miRNAs are small molecules of about 23 or less ribonucleotides which are cleaved from larger fragments of RNA transcripts encoded from within genes, or large tracts of DNA between the genes. miRNAs are a special group of non-coding RNAs which work by blocking (silencing) gene expression. More than 400 have been discovered in the human genome so far and some appear to regulate many genes, as in embryonic development, e.g. cell growth, apoptosis. A cancer-causing miRNA was identified and called *onco-miRNA*. miRNA expression in particular cancers can now be quantified.

Single stranded RNAs consisting of 100s of bases fold back on themselves and undergo complementary base-pairing to form double-stranded “short-hairpin” RNA (shRNA). The loop of the hair-pin is excised by endonucleases, e.g. *Dicer*, to produce multiple functional miRNAs.

Many of the above RNAs for specific purposes are available commercially from several firms in high purity because they are synthetic. These firms will also synthesise particular RNAi’s to suit one’s requirements.

**Subtilisin (from *Bacillus subtilis*)** [9014-01-1]  $M_r$  27,000 (sedimentation equilibrium) [EC 3.4.21.62]. This alkaline protease is purified 211-fold by affinity chromatography using 4-(4-aminophenylazo)phenylarsonic acid complex to activated CH-Sepharose 4B. It is inhibited by 2-phenylethane boronic acid, PMSF, 3,4-dichloroisocoumarin, acetone and benzamide. [Chandraskaren & Dhar *Anal Biochem* 150 141 1985, Schomburg & Schomburg *Springer Handbook of Enzymes* 2nd Edn vol 7 p 286 2002.]

**Synexin (from bovine liver)**  $M$  47,000 Da. This Ca binding protein is purified by  $(NH_4)_2SO_4$  precipitation, then by a specific pH step elution from a chromatofocusing medium in the absence of ampholytes. The pI is 7.5. [Scott et al. *Anal Biochem* 149 163 1985.]

**Thrombin (from bovine blood plasma)** [9002-04-4]  $M_r$  32,600 [EC 3.4.4.13]. Thrombin is purified by chromatography on a DEAE-cellulose column, while eluting with 0.1M NaCl, pH 7.0, followed by chromatography on Sephadex G-200. The final preparation is free from plasminogen and plasmin. [Yin & Wessler *J Biol Chem* 243 112 1968.] Thrombin from bovine blood is also purified by chromatography using *p*-chlorobenzylamino- $\epsilon$ -aminocaproyl agarose, and gel filtration through Sephadex G-25. [Thompson & Davie *Biochim Biophys Acta* 250 210 1971.]

Thrombin from various species was purified initially by precipitation of impurities with Rivanol (see [6402-23-9]). [Miller *Nature* 184 450 1959.]

**Tissue inhibitor of metalloproteins (TIMP, from human blood plasma),**  $M_r$  ~30,000. These are purified by an [anti-human amniotic fluid-TIMP]-Sepharose immuno-affinity column and eluted with 50mM glycine/HCl

pH 3.0 buffer that is 0.5M in NaCl, and followed by gel filtration [Cawston et al. *Biochem J* **238** 677 1986].

**Transferrin (from human or bovine serum)** [11096-37-0]  $M_r$  ~80,000. This transferrin is purified by affinity chromatography on phenyl-boronate agarose followed by DEAE-Sepharcel chromatography. The product is free from haemopexin. [Cook et al. *Anal Biochem* **149** 349 1985, Aisen & Listowsky *Ann Rev Biochem* **49** 357 1980.]

**Trehalase (from kidney cortex,  $\alpha,\alpha$ -trehalose glycohydrolase)** [9025-52-9]  $M_r$  ~80,000 [EC 3.2.1.28]. This trehalase is purified by solubilising in Triton X-100 and sodium deoxycholate, and submitting to gel filtration, ion-exchange chromatography, conA-Sepharose chromatography, phenyl-Sepharose CL-4B hydrophobic interaction chromatography, Tris-Sepharose 6B affinity and hydrolyapatite chromatography. Activity is increased 3000-fold. [Yoneyama *Arch Biochem Biophys* **255** 168 1987.]

**T4-RNA ligase (from bacteriophage-infected *E. coli*)**  $M_r$  43,500, [EC 6.5.1.3 for RNA lyase]. This ligase is purified by differential centrifugation and separation on a Sephadex A-25 column, then through hydroxylapatite and DEAE-glycerol using Aff-Gel Blue to remove DNAase activity. (Greater than 90% of the protein in the enzyme preparation migrated as a single band on gradient polyacrylamide gels containing SDS during electrophoresis.) [McCoy et al. *Biochim Biophys Acta* **562** 149 1979.]

**Ubiquinol-cytochrome c reductase (from beef heart mitochondria)** [9027-03-6] [EC 1.10.2.2]. This reductase is purified by solubilising the crude enzyme with Triton X-100, followed by hydroxylapatite and gel chromatography. The minimum unit contains nine polypeptide subunits of  $M_r$  6000 – 49,000 kD. [Engel et al. *Biochim Biophys Acta* **592** 211 1980.]

**Uridine 5'-diphosphoglucose pyrophosphorylase (from rabbit skeletal muscle)** [9029-22-6]  $M_r$  350,000 [EC 2.7.7.9]. The pyrophosphorylase is purified by two hydrophobic chromatographic steps and gel filtration. [Bergamini et al. *Anal Biochem* **143** 35 1984.] It is also purified from calf liver by  $(\text{NH}_4)_2\text{SO}_4$  (40-58%) precipitation,  $\text{Ca}_3(\text{PO}_4)_2$  gel filtration, DEAE-cellulose chromatography and recrystallisation [by dialysis against concentrations of  $(\text{NH}_4)_2\text{SO}_4$  (from 10%) in 0.02M TEA, at 2.5% increments, until 20%  $(\text{NH}_4)_2\text{SO}_4$  when it crystallises out [Hansen et al. *Methods Enzymol* **8** 248 1966].

**Urokinase (from human urine)** [9039-53-6]  $M_r$  53,000 [EC 3.4.21.31]. Crystallisation of this enzyme is induced at pH 5.0 to 5.3 (4°) by careful addition of NaCl with gentle stirring until the solution becomes turbid (silky sheen). The NaCl concentration is increased gradually (over several days) until 98% of saturation is achieved whereby urokinase crystallises out as colourless thin brittle plates. It can be similarly recrystallised to maximum specific activity [104K CTA units/mg of protein (Sherry et al. *J Lab Clin Med* **64** 145 1964)]. [Lesuk et al. *Science* **147** 880 1965, NMR: Bogusky et al. *Biochemistry* **28** 6728 1989.] It is a plasminogen activator [Gold et al. *Biochem J* **262** 1989, de Bock & Wang *Med Res Rev* **24**(1) 13 2004].

**Xylanase (from *Streptomyces lividans*)** [37278-89-0]  $M_r$  43,000 [EC 3.2.1.8]. This xylanase is purified by anion-exchange chromatography on an Accell QMA column and finally by HPLC using a ProteinPak DEAE 5PW anion-exchange column. Solutions are stored frozen at -70°. [Morosoli et al. *Biochem J* **239** 587 1986, Wong et al. *Microbiol Rev* **52** 305 1988.]

## CARBOHYDRATES

This section includes natural and synthetic carbohydrates and glycosides.

### Z-O-(2-Acetamido-2-deoxy-D-glycopyranosylideneamino)-N-phenyl-carbamate

(PUGNAC) [132063-05-9] M 335.3, m 171-174° (dec), 174-180°(dec),  $[\alpha]_D^{20} +67.5^\circ$  (c 0.2, MeOH). Purify PUGNAC by flash chromatography (silica gel and elute with EtOAc/hexane 3:2), evaporate, and the foam is recrystallised from EtOAc/MeOH. The purity is checked by TLC on Merck SiO<sub>2</sub> gel 60 F<sub>254</sub> and the spots are detected by spraying with 0.025M I<sub>2</sub> in 10% aqueous H<sub>2</sub>SO<sub>4</sub> and heating at 200°. It has R<sub>F</sub> 0.21. The acetate is hydrolysed with NH<sub>3</sub>/MeOH. [*Helv Chim Acta* 68 2254 1985, *Helv Chim Acta* 73 1918 1990.]

**Acetobromo- $\alpha$ -D-galactose** [3068-32-4] M 411.2, m 87°,  $[\alpha]_{546}^{20} +255^\circ$ ,  $[\alpha]_D^{20} +210^\circ$  (c 3, CHCl<sub>3</sub>). Purify acetobromo- $\alpha$ -D-galactose as for the glucose analogue (next entry). If the compound melts lower than 87° or is highly coloured, then dissolve it in CHCl<sub>3</sub> (ca 3 volumes) and extract with H<sub>2</sub>O (2 volumes), 5% aqueous NaHCO<sub>3</sub>, and again with H<sub>2</sub>O and dry it over Na<sub>2</sub>SO<sub>4</sub>. Filter and evaporate it in a vacuum. The partially crystalline solid or syrup is dissolved in dry Et<sub>2</sub>O (must be very dry) and recrystallised by adding petroleum ether (b 40-60°) to give a white product. [McKellan & Horecker *Biochemical Preparations* 11 111 1960, *Beilstein* 17/6 V 369.]

**Acetobromo- $\alpha$ -D-glucose ( $\alpha$ -acetobromoglucose 2,3,4,6-tetraacetyl- $\alpha$ -D-glycopyranosyl bromide)** [572-09-8] M 411.2, m 87-88°, 88-89°,  $[\alpha]_{546}^{20} +230^\circ$ ,  $[\alpha]_D^{20} +195^\circ$  (c 3, CHCl<sub>3</sub>). If nicely crystalline, recrystallise it from Et<sub>2</sub>O/pentane or petroleum ether (b 40-60°). Alternatively, dissolve it in diisopropyl ether (dried over CaCl<sub>2</sub> for 24 hours, then over P<sub>2</sub>O<sub>5</sub> for 24 hours) by shaking and warming (for as short a period as possible), and filter warm. Cool to ca 45°, then slowly to room temperature and finally at 5° for more than 2 hours. Collect the solid, wash it with cold dry diisopropyl ether and dry it in a vacuum over Ca(OH)<sub>2</sub> and NaOH. Store it dry in a desiccator in the dark. Solutions can be stabilised with 2% CaCO<sub>3</sub>. [Redemann & Niemann *Org Synth* 65 236 1987, Coll Vol III 11 1955, *Beilstein* 17/6 V 368.]

**$\alpha$ -Acetyldigitoxin** [1111-39-3] M 807.0, m 153-180° (monohydrate, dec), 217-221°, 258°(dec),  $[\alpha]_D^{20} +5.0^\circ$  (c 0.7, pyridine), +24.5° (c 1, MeOH).  $\alpha$ -Acetyldigitoxin is obtained from the commercial mixture [ $\alpha$ : $\beta$  (2:1)] [25395-32-8]. The  $\alpha$ -form is obtained from the  $\beta$ -form by heating in anhydrous or aqueous organic solvent (e.g. aqueous MeOH) at pH 3.5—8. It crystallises from MeOH as plates, CHCl<sub>3</sub>/Et<sub>2</sub>O or Me<sub>2</sub>CO/Et<sub>2</sub>O. At 20° 1g dissolves in 16ml of MeOH, 66ml of Me<sub>2</sub>CO and ~880ml of EtOAc. [Stoll & Kreis *Helv Chim Acta* 35 1318, 1322 1952, *Beilstein* 18 III/IV 1479.] It is a cardenolide.

**$\beta$ -Acetyldigitoxin** [1264-51-3] M 807.0, m 225°,  $[\alpha]_D^{20} +18.7^\circ$  (c 0.7, pyridine), +24.3° (c 0.7, MeOH).  $\beta$ -Acetyldigitoxin crystallises from MeOH as a methanolate which loses MeOH in a vacuum desiccator. The solubility at room temperature is: H<sub>2</sub>O (0.0005%), EtOAc (0.4%), MeOH (0.6%), Me<sub>2</sub>CO (1.6%) and CHCl<sub>3</sub>(12%). [Stoll & Renz *Helv Chim Acta* 35 1310 1952, Kreis US 2776963, 1957 to Sandoz; *Beilstein* 18 III/IV 1479]. It is a cardenolide.

**N-Acetyl-D-lactosamine [2-acetyl-amino-O- $\beta$ -D-lactopyranosyl-2-deoxy-D-glucose]** [32181-59-2] M 383.4, m 169-171°, 170-171°,  $[\alpha]_D^{18} +51.5^\circ \rightarrow +28.8^\circ$  (in 3 hours, c 1, H<sub>2</sub>O). Purify N-acetyl-D-lactosamine by recrystallisation from MeOH (with 1 mol of MeOH) or from H<sub>2</sub>O. It is available commercially as a solution of 0.5g/ml of H<sub>2</sub>O. [Zilliken *J Biol Chem* 271 181 1955, *Beilstein* 17 IV 3452.]

**1-O-Acetyl-2,3,5-tri-O-benzoyl- $\beta$ -D-ribofuranose** [6974-32-9] M 504.5, m 128-130°, 130-131°, 131-132°,  $[\alpha]_D^{20} +44.2^\circ$  (c 1, CHCl<sub>3</sub>). Recrystallise it from EtOH or isoPrOH. [*Helv Chim Acta* 42 1171 1959, NMR: *J Org Chem* 33 1799 1968, IR: *Chem Pharm Bull Jpn* 11 188 1963, *Beilstein* 17/6 V 213.]

**N-Acetyl muramic acid [NAM]** See “Miscellaneous” in this Chapter.

**Alginic acid** [9005-32-7] **M 48,000-186000**. To 5g of acid in 550ml water containing 2.8g  $\text{KHCO}_3$  are added 0.3ml of acetic acid and 5g potassium acetate. EtOH is added to make the solution to 25% (v/v) in EtOH, and any insoluble material is discarded. Further addition of EtOH, to 37% (v/v), precipitated alginic acid. Collect the acid and dry it *in vacuo*. [Pal & Schubert *J Am Chem Soc* **84** 4384 1962.]

**Aloin [10-glucopyranosyl-1,8-dihydroxy-3-(hydroxymethyl)-9(10H)anthracenone, Barbaloin]** [8015-61-0, 5133-19-7, 1415-73-2] **M 418.4, m 148-148.5<sup>o</sup>, 148-150<sup>o</sup>**. Aloin forms the *monohydrate* as lemon yellow crystals on crystallisation from  $\text{H}_2\text{O}$  (450g/1.5l) which has a lower **m** ~70-80<sup>o</sup> than the anhydrous substance. [Cahn & Simonsen *J Chem Soc* 2573 1932, Hay & Haynes *J Chem Soc* 3141 1956.]

**D-Altrose** [1990-29-0] **M 180.2, m 103-105<sup>o</sup>,  $[\alpha]_{546} +35^{\circ}$  (c 7.6,  $\text{H}_2\text{O}$ )**. Crystallise D-altrose from aqueous EtOH. If it is obtained by the hydrolysis of the acetate, then it may contain sodium and acetate ions. Ions are best removed by dissolving in  $\text{H}_2\text{O}$ , passing through suitable columns of ion-exchange resins, e.g. Amberlite IR-120 and Duolite A, and concentrating in a vacuum to a syrup. This is dissolved in MeOH, filtered and evaporated in a vacuum desiccator over granular  $\text{CaCl}_2$ . The thick syrup is inoculated with seed crystals, stirred, and before it sets to a magma of crystals, transfer the crystals with MeOH to a Büchner funnel. Recrystallise them in the same way.  $\beta$ -D-Altrofuranside has initial  $[\alpha]_{20}^{\circ} \sim -69^{\circ}$  (c 4,  $\text{H}_2\text{O}$ ) which mutarotates to  $+33^{\circ}$ . [Richtmeyer *Methods in Carbohydrate Chemistry* **I** 107 Academic Press 1962, Beilstein **1** IV 4301, see Angyal *Adv Carbohydrate Chem Biochem* **42** 15 1984 for ratio of anomers in solution.]

**D-Amygdalin (D-mandelonitrile- $\beta$ -gentiobioside)** [29883-15-6] **M 457.4, m 214-216<sup>o</sup>,  $[\alpha]_{\text{D}}^{\circ} -38^{\circ}$  (c 1.2,  $\text{H}_2\text{O}$ )**. D-Amygdalin recrystallises from water as the *trihydrate*, or from EtOH. It is present in bitter almonds. [Smith *Chem Ber* **64** 1115 1931, Beilstein **17/8** V 188.]

**Amylose** [9005-82-7]  $(\text{C}_6\text{H}_{10}\text{O}_5)_n$  (*for use in iodine complex formation*). Amylopectin is removed from impure amylose by dispersing in aqueous 15% pyridine at 80-90<sup>o</sup> (concentration 0.6-0.7%) and allowing the solution to stand at 44-45<sup>o</sup> for 7 days. The precipitate is re-dispersed and recrystallises during 5 days. After a further dispersion in 15% pyridine, it is cooled to 45<sup>o</sup>, allowed to stand at this temperature for 12 hours, then cooled to 25<sup>o</sup> and left for a further 10 hours. The combined precipitates are dispersed in warm water, precipitated with EtOH, washed with absolute EtOH, and dried *in vacuo* [Foster & Paschall *J Am Chem Soc* **75** 1181 1953.]

**1,6-Anhydro- $\beta$ -D-glucose (levoglucosan,  $\beta$ -glucosan,  $\beta$ -D-pyranose-form)** [498-07-7] **M 162.1, m 177<sup>o</sup>, 179-180<sup>o</sup>, 182-184<sup>o</sup>, 184<sup>o</sup>,  $[\alpha]_{\text{D}} -66.3^{\circ}$  (c 2,  $\text{H}_2\text{O}$ )**. Levoglucosan is readily formed by pyrolysis (dry distillation) of starch, ash-free amylose, cellulose and  $\beta$ -D-glucose in a vacuum. Best results are obtained when the acidity of the distillate is least. The product is dissolved in MeOH, decolourised with charcoal and cooled to give the glucosan in good yields. Other  $\beta$ -D-glucosides such as Arbutin, Salicin and Phloridzin also yield levoglucosan, and is evidence for the  $\beta$ - linkage configuration because  $\alpha$ -glucose and  $\alpha$ -glucosides require more drastic conditions for pyrolysis, i.e. longer heating and higher temperatures since the  $\alpha$ -orientation at C-1 needs to be inverted to the  $\beta$ -orientation [Karrer *Helv Chim Acta* **3** 258 1920, Pictet & Cramer *Helv Chim Acta* **3** 640 1920, Ruchel & Schuerch *J Org Chem* **31** 2233 1966, Irvine & Oldham *J Chem Soc* 2903 1925, Carvalho et al. *J Am Chem Soc* **81** 4054 1959].

Another ready source of levoglucosan is  $\beta$ -phenylglucoside (2.0, see [1464-44-4]) which is dissolved 1.3N aqueous KOH (100ml), heated on boiling water bath for 9 hours, while monitoring the specific optical rotation until it is constant ( $-64.4^{\circ}$ ). The clear mixture is cooled, neutralised (to methyl orange, the slight colour of the cooled solution does not interfere) with 3N sulfuric acid, evaporated to dryness *in vacuo*, extracted from the  $\text{K}_2\text{SO}_4$  residue with warm EtOH, evaporated and the residue is recrystallised from EtOH (10ml) to give *levoglucosan*, m 179-180<sup>o</sup>,  $[\alpha]_{\text{D}} -66.3^{\circ}$  (c 2,  $\text{H}_2\text{O}$ ), in 88% yield. Other  $\beta$ -phenylglyco-pyrano-sides give anhydro-glycosans, but the  $\alpha$ -linked sugars require much longer heating (~20 hours or weeks) to give the anhydro-sugars. This alkaline reaction has been studied in detail, and similarly  $\beta$ -phenylgalactoside provided a 91% yield of **1,6-anhydro- $\beta$ -D-galactose (anhydroglactosan)** [644-76-8] **M 162.1, with m 226<sup>o</sup>** (plates from 50% aqueous EtOH) and  $[\alpha]_{\text{D}} -22.0^{\circ}$  (c 2,  $\text{H}_2\text{O}$ ). [Montgomery et al. *J Am Chem Soc* **65** 3 1943.]

Levoglucosan is one of the more important anhydro-sugars as it, and its 2-, 3- and 4-substituted derivatives, are readily hydrolysed by aqueous acid to give glucose and the respective derivatives [Peat *Adv Carbohydrate*

*Chem* **2** 29 1946]. Also at elevated temperature, in the presence of zinc dust, ZnCl<sub>2</sub>, or platinum black, levoglucosan polymerises to form a variety of oligomers described as *dextrans*, and these can be commonly 20,000 to 50,000 mers (but 300,000 mers in the extreme); and have found use in chromatographic purification of large molecules, e.g. proteins. The polymerisation process has been studied in detail [Carvalho et al. *J Am Chem Soc* **81** 4054 1959, Irvine & Oldham *J Chem Soc* 2903 1925, Ruchel & Schuerch *J Org Chem* **31** 2233 1966, see also Pictet & Sarasin *Helv Chim Acta* **1** 87 1918 and later papers]. The 2-, 3- and 4- OH groups are readily derivatised in the usual way, and the *triacetyl* derivative has [13242-55-2] M 288.2, **m** 109.5-110.5°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -50.8° (CHCl<sub>3</sub>), the *trifluoroacetyl* derivative has M 450.2, **m** 63.5-65.0°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -39.9° (**c** 2.1, CHCl<sub>3</sub>), the *tribenzoyl* derivative has [13567-05-7] M 474.4, **m** 202-203°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -36.4° (CHCl<sub>3</sub>), the *tri-O-methyl* derivative has M 204.2, **m** 57.5-58.5°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -63.7° (**c** 2, H<sub>2</sub>O), and the *tri-O-ethyl* derivative has M 246.3, **b** 90°/0.60mm, **d**<sup>28</sup> 1.081, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -51.5° (**c** 2.8, CHCl<sub>3</sub>), the *tri-O-benzyl* derivative has M 432.5, **m** 89.5-90.5°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -30.8° (**c** 2.7, CHCl<sub>3</sub>), and the *tri-O-methylsilyl* derivative has M 378.7, **b** 87-87.5°/0.05mm, [ $\alpha$ ]<sub>D</sub><sup>30</sup> -33.8° (**c** 3.6, CHCl<sub>3</sub>); all of which have been subjected to the polymerisation process with the “*ether*” derivatives being more successful than the “*ester*” derivatives in producing high polymers. [Ruckel & Scheurch *J Org Chem* **31** 2233 1966]. The <sup>1</sup>H NMR of levoglucosan and its derivatives in DMSO-*d*<sub>6</sub> and D<sub>2</sub>O have been reported in detail [Woolwage & Seib *J Chem Soc (C)* 3143 1971]. [*Beilstein* **19** H 894, **19** I 452, **19** III/V 1181, **19** V 498.]

**β-L-(+)-Arabinose (natural)** [87-72-9] M 150.1, **m** 158°, [ $\alpha$ ]<sub>D</sub> +104° (**c** 4, H<sub>2</sub>O after 24h), pK<sup>17</sup> 12.4. β-L-(+)-Arabinose is recrystallised slowly, twice, from 80% aqueous EtOH, then dried under vacuum over P<sub>2</sub>O<sub>5</sub>. It can also be purified by heating the arabinose (200g) with glacial acetic acid (300ml) on a boiling water bath for 45 minutes, cooling, filtering, washing with 95% EtOH (500ml) in four portions and drying at 56-60° over P<sub>2</sub>O<sub>5</sub>. It has been recrystallised from 5 times its weight of 76% EtOH using charcoal (10g) to yield 127g, **m** 155-157°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +190.6° and mutarotating to +104° (**c** 4, H<sub>2</sub>O). [Anderson & Sands *Org Synth Coll Vol I* 67 1941, Wolfrom & Christian *J Am Chem Soc* **48** 3172 1926, *Beilstein* **1** IV 4217.]

**D-(-)-Arabinose** [10323-20-3, 28697-53-2 (*pyranoside*)] M 150.1, **m** 164°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -104.5° (**c** 1, H<sub>2</sub>O), [ $\alpha$ ]<sub>546</sub> -123° (**c** 10, H<sub>2</sub>O after 24 hours), pK<sup>25</sup> 12.54. Crystallise D-(-)-arabinose three times from EtOH, dry it *in vacuo* at 60° for 24 hours and store it in a vacuum desiccator. It also crystallises from a mixture of H<sub>2</sub>O/MeOH/EtOH with **m** 158-160°, or H<sub>2</sub>O with **m** 160-161°. [Whistler & Schweiger *J Am Chem Soc* **81** 5190 1959, Fletcher et al. *J Am Chem Soc* **72** 4546 1950, *Beilstein* **1** IV 4215, see Angyal *Adv Carbohydrate Chem Biochem* **42** 15 1984 for ratio of anomers in solution.]

**D-Arabitol (D-arabinitol)** [488-82-4] M 152.2, **m** 103°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +13° (**c** 5, 8% borax solution). This pentol, which occurs in lichens and fungi, is purified by recrystallisation from 90% EtOH or MeOH. [Ashina & Yamagita *Chem Ber* **67** 801 1934, derivarives: Nakagawa et al. *Bull Chem Soc Jpn* **40** 2150 1967, Prince & Reichstein *Helv Chim Acta* **20** 101 1937, Hough & Theobald *Methods in Carbohydrate Chemistry* **I** 94 1962, Academic Press, *Beilstein* **1** IV 2832.]

**L-Arabitol (L-arabinitol)** [7643-75-6] M 152.2, **m** 102°, 103°, [ $\alpha$ ]<sub>546</sub> -13° (**c** 5, 8% borax solution). This pentol, which occurs in the urine of pentosuric subjects, is purified by recrystallisation from 90% EtOH or MeOH. It has a higher rotation in the presence of molybdate: [ $\alpha$ ]<sub>D</sub><sup>20</sup> -130° (**c** 0.16, acidified molybdate) [Richtmeyer & Hudson *J Am Chem Soc* **73** 2249 1957]. [Gätzi & Reichstein *Helv Chim Acta* **21** 197 1938, *Beilstein* **1** IV 2832.]

**DL-Arabitol (DL-arabinitol)** [2152-56-9] M 152.2, **m** 105-106°. This synthetic arabitol is purified by recrystallisation from 90% EtOH, MeOH or EtOH/Me<sub>2</sub>CO. [Raphael *J Chem Soc, Supplement*, 48 1949, Ashira & Yamagita *Chem Ber* **67** 802 1934, *Beilstein* **1** IV 2832.]

**p-Arbutin (p-hydroquinone-O-β-D-glucopyranoside)** [497-76-7] M 272.3, **m** 195-198°, 199°, 200°(sintering at 163-164°), [ $\alpha$ ]<sub>D</sub><sup>20</sup> -65° (**c** 2, H<sub>2</sub>O), pK<sub>Est</sub> ~10.0. The glycoside from *Protea exima* is purified by recrystallisation from H<sub>2</sub>O or moist EtOAc (as *monohydrate*), after chromatography through silica Gel using EtOAc/MeOH. Crystallisation from EtOH/CHCl<sub>3</sub> gives crystals **m** 199-200° with intermediate melting at 164° and resolidifying. The *pentaacetate* crystallises from EtOH in fine needles with **m** 145-146°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -28.2°

(c 2, Me<sub>2</sub>CO). [Robinson & Waters *J Chem Soc* 2729 1930, IR, NMR, MS: Perold et al. *J Chem Soc, Perkin Trans 1* 239 1979, *Beilstein* 17/7 V 110.]

**L(+)-Ascorbic acid** [50-81-7] **M 176.1, m 193°(dec), [α]<sub>546</sub> +23° (c 10, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 4.04, pK<sub>2</sub><sup>25</sup> 11.34.** Crystallise it from MeOH/Et<sub>2</sub>O/petroleum ether [Herbert et al. *J Chem Soc* 1270 1933]. [*Beilstein* 18/5 V 26.]

**6-Bromo-2-naphthyl-α-D-galactopyranoside** [25997-59-5] **M 385.2, m 178-180°, 224-226°, 225°, [α]<sub>D</sub><sup>28</sup> +60° (c 1.2, pyridine).** It is prepared from penta-*O*-acetyl-D-galactoside, 6-bromo-2-naphthol and ZnCl<sub>2</sub>. The resulting tetra-acetate (2g) is hydrolysed by dissolving in 0.3N KOH (100ml) and heating until the solution is clear, then filtering and cooling to give colourless crystals of the α-isomer which are collected and recrystallised twice from hot MeOH. The high specific rotation is characteristic of the α-isomer. The *tetraacetate* has **m 155-156°, [α]<sub>D</sub><sup>20</sup> +60° (c 1, CHCl<sub>3</sub>)** [Dey & Pridham *Biochem J* 115 47 1969] [reported **m 75-85°, [α]<sub>D</sub><sup>24</sup> +94° (c 1.3, dioxane)**, Monis et al. *J Histochem Cytochem* 11 653 1963]. [*Beilstein* 17 IV 2972.]

**λ-Carrageenan** [9064-57-7 (κ), 9000-07-1 (κ + little of λ)]. This D-galactose-anhydro-D or L-galactoside polysaccharide is precipitated from 4g of Carrageenan in 600ml of water containing 12g of KOAc by addition of EtOH. Collect the fraction that precipitates between 30 and 45% (v/v) of EtOH and dry them *in vacuo*. [Pal & Schubert *J Am Chem Soc* 84 4384 1962.]

**Colchicoside** [7S-7N-[3-(β-D-glucopyranosyloxy)-1,2,10-tetramethoxy-9-oxobenzo[*a*]-heptalen-7-yl-(7S)]-acetamide, 7-glucosylcolchicine] [477-29-2] **M 547.5, m 216-218° (block), [α]<sub>D</sub><sup>15</sup> -128.5° (c 1, CHCl<sub>3</sub>), [α]<sub>D</sub><sup>15</sup> -231° (c 1, MeOH), [α]<sub>D</sub><sup>15</sup> -355° (c 1, H<sub>2</sub>O).** Purify colchicoside by chromatography through alumina and eluting with CHCl<sub>3</sub>, then recrystallising from Me<sub>2</sub>CO, **m 275-280°** (after releasing solvent at 180°). It also crystallises from EtOH; see colchicine [64-86-8]. The *tetraacetate* has **m 175-177°, [α]<sub>D</sub><sup>15</sup> -57° (c 1, CHCl<sub>3</sub>).** [Bellet et al. *Ann Pharm Fr* 10 241 1952, *Chem Abstr* 47 3323 1953, *Beilstein* 14 IV 946 for colchicine.]

**Convallatoxin (α-cardenolide mannoside)** [508-75-8] **M 550.6, m 238-239°, 238-241° (dec), [α]<sub>D</sub><sup>20</sup> -9.4° (c 0.7, dioxane), [α]<sub>D</sub><sup>16</sup> -1.0 ±3° (c 0.7, EtOH).** Crystallise convallatoxin from EtOAc, CHCl<sub>3</sub>/EtOH (9:1) or MeOH/Et<sub>2</sub>O. The *tetraacetate* has **m 238-242°** (from MeOH/Et<sub>2</sub>O), [α]<sub>D</sub><sup>25</sup> -5° (CHCl<sub>3</sub>). [Reyle et al. *Helv Chim Acta* 33 1541 1950, Fieser & Jacobson *J Am Chem Soc* 59 2335 1937 *Beilstein* 18 III/IV 3142.]

**α-Cyclodextrin (H<sub>2</sub>O)** [10016-20-3] **M 972.9, m >280°(dec), [α]<sub>546</sub> +175° (c 10, H<sub>2</sub>O).** Recrystallise α-cyclodextrin from 60% aqueous EtOH, then twice from water, and dry it for 12 hours in a vacuum at 80°. It is also purified by precipitation from water with 1,1,2-trichloroethylene. The precipitate is collected, washed and resuspended in water. This is boiled to steam distil the trichloroethylene. The solution is then freeze-dried to recover the cyclodextrin. [Armstrong et al. *J Am Chem Soc* 108 1418 1986]. [*Beilstein* 19/12 V 789.]

**β-Cyclodextrin (H<sub>2</sub>O)** [7585-39-9, 68168-23-0] **M 1135.0, m >300°(dec), [α]<sub>546</sub> +170° (c 10, H<sub>2</sub>O).** Recrystallise β-cyclodextrin from water and dry it for 12 hours in a vacuum at 110°, or 24 hours in a vacuum at 70°. The purity is assessed by TLC on cellulose containing a fluorescent indicator. [Taguchi, *J Am Chem Soc* 108 2705 1986, Tabushi et al. *J Am Chem Soc* 108 4514 1986, Orstam & Ross *J Phys Chem* 91 2739 1987]. [*Beilstein* 19 IV 6287, 19/12 V 801.]

***N*-Decanoyl-*N*-methylglucamine (Mega-10, *N*-D-glucidyl-*N*-methyl deconamide)** [85261-20-7] **M 349.5, m 91-93°, 92°.** Possible impurities are decanoic acid and *N*-methylglycamine. The former is removed by grinding the solid with Et<sub>2</sub>O and then with petroleum ether and drying over P<sub>2</sub>O<sub>5</sub>. It is twice recrystallised from MeOH/Et<sub>2</sub>O by dissolving in the minimum volume of MeOH, adding Et<sub>2</sub>O and drying in a vacuum. To remove the glycamine, the solid (800mg) is dissolved in hot H<sub>2</sub>O (10ml) and set aside. Mega-10 crystallises as colourless needles. These are filtered off and dried in a vacuum to constant weight. It is a good non-ionic non-hygroscopic detergent with a critical micelle concentration (CMC) of 7.4mM (0.26%) in 0.1M Tris-HCl pH 7.4

at 25°. [Hildreth *Biochem J* **207** 363 1982.]

**Dehydro-L(+)-ascorbic acid** [490-83-5] **M 174.1, m 196°(dec),  $[\alpha]_{546}^{20} +42.5^\circ$  (c 1, H<sub>2</sub>O), pK<sup>25</sup> 3.90.** Crystallise dehydro-L(+)-ascorbic acid from MeOH. The *anhydrous acid* is formed by heating it in a vacuum at 100°/1 hour to give a crisp glassy product which when shaken with absolute EtOH and then kept at 0° for 2 days gives microcrystals of the anhydrous acid. This is then washed with absolute EtOH and dried in a vacuum. It has **m 225°(dec)** and is stable in acidic solution but decomposes rapidly in alkaline solution. A 1% solution of the anhydrous acid when dissolved in phthalate/HCl buffer pH 3.5 at 60° and cooled to 20° has  $[\alpha]_{\text{D}}^{20} +56^\circ$  (0 minutes), +53.5°(2 hours), +19°(3 days), -2°(5 days) and -6°(6 days); then it becomes orange in colour. A freshly prepared 1% solution in H<sub>2</sub>O has  $[\alpha]_{\text{D}}^{20} +50^\circ$  (0 minutes), +44°(2 hours), +16°(3 days) and 0°(5 days). [Herbert et al. *J Chem Soc* 1270 1933, Kenyon et al. *J Chem Soc* 158 1948, *Beilstein* **18/5** V 411.]

**2-Deoxy-D-allose (2-deoxy-D-ribo-hexose)** [6605-21-6] **M 164.2, m 140-142°,  $[\alpha]_{\text{D}}^{24} +57.5^\circ$  (c 1.2, H<sub>2</sub>O).** Purify 2-deoxy-D-allose by two recrystallisations from absolute EtOH. The *p-nitrophenylhydrazone* has **m 61-62°,  $[\alpha]_{\text{D}}^{16.5} -55^\circ$  (MeOH).** An equilibrium solution at 31° in D<sub>2</sub>O contains 15%  $\alpha$ -pyranose, 58%  $\beta$ -pyranose, 12%  $\alpha$ -furanose and 15%  $\beta$ -furanose forms as estimated by <sup>1</sup>HNMR spectroscopy. [see Angyal *Adv Carbohydrate Chem Biochem* **42** 15 1984 for ratio of anomers in solution, Zorbach & Ollapally *J Org Chem* **129** 1790 1964.] [*Beilstein* **1** IV 4283.]

**2-Deoxy- $\beta$ -D-galactose (2-deoxy-D-lyxo-hexose)** [1949-89-9] **M 164.2, m 110°, 120-121°, (126-128°),  $[\alpha]_{\text{D}}^{20} +60^\circ$  (c 2, H<sub>2</sub>O, 2 hours).** Crystallise 2-deoxy- $\beta$ -D-galactose from MeOH or diethyl ether. The *aniline derivative* has **m 142-143°,  $[\alpha]_{\text{D}}^{16.5} -149^\circ$  (c 0.8, pyridine).** [Overend et al. *J Chem Soc* 671, 675 1950 and 992 1951.] A 30% equilibrium solution at 31° in D<sub>2</sub>O contains 40%  $\alpha$ -pyranose, 44%  $\beta$ -pyranose, 8%  $\alpha$ -furanose and 8%  $\beta$ -furanose forms as estimated by <sup>1</sup>HNMR spectroscopy [Angyal & Pickles *Aust J Chem* **25** 1711 1972]. [*Beilstein* **1** IV 4283.]

**2-Deoxy- $\alpha$ -D-glucose (2-deoxy-D-arabino-hexose)** [154-17-6] **M 164.2, m 148-151°,  $[\alpha]_{\text{D}}^{20} +46^\circ$  (c 0.5, H<sub>2</sub>O after 45 hours).** Crystallise 2-deoxy- $\alpha$ -D-glucose from MeOH/Me<sub>2</sub>CO, Me<sub>2</sub>CO or butanone to give a mixture of  $\alpha$ - and  $\beta$ - anomers, **m 142-144°,  $[\alpha]_{\text{D}}^{18} +38^\circ$  (35 minutes) to +46° (c 0.5, H<sub>2</sub>O).** Recrystallisation from *iso*PrOH gives mainly the  $\alpha$ -anomer **m 134-136°,  $[\alpha]_{\text{D}}^{25} +156^\circ$  to +103° (c 0.9, pyridine).** <sup>1</sup>H NMR studies showed that at 44° in D<sub>2</sub>O the solution contained 36% of  $\alpha$ -pyranose and 64% of  $\beta$ -pyranose sugar, but furanose structures were undetectable. [Snowden & Fischer *J Am Chem Soc* **69** 1048 1947, derivatives: Bollinger & Schmidt *Helv Chim Acta* **34** 989 1951; see Angyal & Pickles *Aust J Chem* **25** 1711 1972 for ratio of isomers in solution, *Beilstein* **1** IV 4282.]

**6-Deoxy-D-glucose (D-quinovose)** [7658-08-4] **M 164.2, m 146°,  $[\alpha]_{\text{D}}^{20} +73^\circ$  (after 5 minutes) and +30° (final, after 3 hours) (c 8.3, H<sub>2</sub>O).** 6-Deoxy-D-glucose is purified by recrystallisation from EtOAc and is soluble in H<sub>2</sub>O, EtOH but insoluble in Et<sub>2</sub>O and Me<sub>2</sub>CO. [Srivastava & Lerner *Carbohydr Res* **64** 263 1978; NMR: Angyal & Pickles *Aust J Chem* **25** 1711 1972, *Beilstein* **1** IV 4260.]

**2-Deoxy- $\beta$ -L-ribose** [18546-37-7] **M 134.1, m 77°, 80°,  $[\alpha]_{\text{D}}^{25} +91.7^\circ$  (c 7, pyridine, +40° final).** Crystallise 2-deoxy- $\beta$ -L-ribose from diethyl ether. It can also be purified by dissolving the ribose (7.3g) in EtOAc (3L) by reflux, decanting from any insoluble material and evaporating at 50°/vacuum to 2.2L, setting aside for 1-2 hours, filtering, and concentrating to 840ml. The ribose separates as compact nodules during 3-5 days at 0° and has **m 87-93°, and after repeated recrystallisations it has m 92-95°.** Mutarotation is as follows:  $[\alpha]_{\text{D}}^{16.5} +80^\circ$  (0 minutes), +71°(6 minutes), +59°(21 minutes) and +59°(41 minutes) (c 1.14, H<sub>2</sub>O);  $[\alpha]_{\text{D}}^{20} +105^\circ$  (0 minutes), +71°(9.5 minutes), +67°(12.5 minutes), +49°(81 minutes), +49°(136 minutes) (c 0.9, MeOH). [Deriaz et al. *J Chem Soc* 1879 1949, *Beilstein* **1** IV 4181.]

**2-Deoxy- $\beta$ -D-ribose** [533-67-5] **M 134.1, m 86-87°, 87-90°,  $[\alpha]_{\text{D}}^{20} -56^\circ$  (c 1, H<sub>2</sub>O after 24 hours).** Dissolve 2-deoxy- $\beta$ -D-ribose in a little H<sub>2</sub>O, evaporate to a syrup (in a vacuum), and seed to crystallise. Triturate the crystals with a little EtOAc containing 5% MeOH, decant and dry in vacuum over P<sub>2</sub>O<sub>5</sub>. It is best purified *via* the *anilide* which separates from a mixture of the ribose (100-125g) in MeOH (100ml) and redistilled aniline (40ml) in a few minutes. After standing for 20 hours at room temperature, it is cooled to 0°, filtered, washed

with 50% aqueous MeOH and Et<sub>2</sub>O followed by recrystallisation from ethylene glycol monomethyl ether. The anilide has **m 172-173°**,  $[\alpha]_D^{25} +46^\circ$  (equilibrium in pyridine). The anilide (5g), benzaldehyde (5ml) and benzoic acid (0.5g) in H<sub>2</sub>O (150ml) are shaken mechanically for 20-24 hours. The aqueous phase is extracted with Et<sub>2</sub>O (3x), decolourised with a little charcoal and evaporated in a vacuum to a syrup. This is dried over P<sub>2</sub>O<sub>5</sub> in high vacuum. The syrupy sugar weighs 3.1g and crystallises in a few days, but more rapidly on seeding. Triturate it with a little EtOAc containing 5% MeOH, decant and dry it over P<sub>2</sub>O<sub>5</sub>. At this stage it has **m 78-82°**,  $[\alpha]_D^{25} -57^\circ$  (c 1, H<sub>2</sub>O final). This is a mixture of  $\alpha$ - and  $\beta$ - anomers. Pure  **$\beta$ -anomer** is obtained by recrystallisation from EtOAc. The  $\beta$ -anomer when recrystallised from EtOAc and isoPrOH has **m 96-98°**,  $[\alpha]_D^{25} -55^\circ$  (c 0.5, H<sub>2</sub>O final). [Sowden *Biochemical Preparations* 5 75 1957.] The mutarotation is as follows:  $[\alpha]_D^{20.5} +96.3^\circ$  (0 minutes),  $-76^\circ$  (33 minutes),  $-56^\circ$  (24 hours) (c 5.8 MeOH). It is moderately hygroscopic and should be kept in a well stoppered bottle. It also crystallises from diethyl ether. [Deriaz et al. *J Chem Soc* 1879 1949, *Beilstein* 1 IV 4181, Hauske & Rapoport *J Org Chem* 44 2472 1979.]

The **1,3,4-tribenzoate ( $\alpha$ - and  $\beta$ -mixture)**, obtained by benzylation, has **m 127°** (after crystallisation from EtOH),  $[\alpha]_D^{25} -65^\circ$  (c 1, CHCl<sub>3</sub>). The crude syrupy mixture in \*C<sub>6</sub>H<sub>6</sub> is applied to an acid-washed Alorco Al<sub>2</sub>O<sub>3</sub> column. Elution with \*C<sub>6</sub>H<sub>6</sub>/hexane (1:1) affords (after crystallisation from MeOH),  **$\beta$ -1,3,4-tri-O-benzoyl-2-deoxy-D-ribose**, **m 159-169°**,  $[\alpha]_D^{20} -196^\circ$  (c 1, CHCl<sub>3</sub>). Further elution with \*C<sub>6</sub>H<sub>6</sub> gives, after recrystallisation from MeOH pure  **$\alpha$ -1,3,4-tri-O-benzoyl-2-deoxy-D-ribose**, **m 151-152°**,  $[\alpha]_D^{20} +41.6^\circ$  (c 0.83, CHCl<sub>3</sub>). [Pedersen et al. *J Am Chem Soc* 82 3425 1960; see Angyal *Adv Carbohydrate Chem Biochem* 42 15 1984 for ratio of anomers and ring forms.]

**Dextran** [9004-54-0] **M<sub>r</sub> 6,000-220,000**. Solutions of dextran keep indefinitely at room temperature if 0.2ml of Roccal (10% alkyl dimethylbenzylammonium chloride) or 2mg phenyl mercuric acetate are added per 100ml solution. This inhibits mould growth. [Scott & Melvin *Anal Biochem* 25 1656 1953.]

**Diacetone-D-glucose (1,2:5,6-di-O-isopropylidene- $\alpha$ -D-glucofuranoside)** [582-52-5] **M 260.3, m 107-110°, 110.5°, 111-113°, 112°**,  $[\alpha]_D^{15} -18.4^\circ$  (c 1, H<sub>2</sub>O). Diacetone-D-glucose crystallises from Et<sub>2</sub>O, (needles), petroleum ether or \*C<sub>6</sub>H<sub>6</sub> and sublimes *in vacuo*. It is soluble in 7 volumes of H<sub>2</sub>O and 200 volumes of petroleum ether at their boiling points. The solubility in H<sub>2</sub>O at 17.5° is 4.3%. It precipitates from aqueous solutions on basification with NaOH. [Schmid & Karrer *Helv Chim Acta* 32 1371 1949, Fischer & Rund *Chem Ber* 49 90, 93 1916, IR: Kuhn *Anal Chem* 22 276 1950, *Beilstein* 19/12 V 318.]

***N,N'*-Diacetylchitobiose (2-acetyl-O<sup>4</sup>-[2-acetylamino-2-deoxy- $\beta$ -D-glucopyranosyl]-2-deoxy-D-glucose)** [35061-50-8] **M 424.4, m 245-247°(dec), 251.5-252.5°, 260-262°**,  $[\alpha]_D^{25} +39.5^\circ$  (extrapolated)  $\rightarrow +18.5^\circ$  (after 60 minutes, c 1, H<sub>2</sub>O). Recrystallise *N,N'*-Diacetylchitobiose from aqueous MeOH or aqueous EtOH/1,2-dimethoxyethane. [Zilliken et al. *J Chem Soc* 77 1296 1955, *Beilstein* 18/11 V 147.]

**1,3,4,6-Di-O-benzylidene-D-mannitol** [28224-73-9] **M 358.4, m 192-195°, 193°**,  $[\alpha]_D^{20} -11.9^\circ$  (c 0.7, Me<sub>2</sub>CO). 1,3,4,6-Di-O-benzylidene-D-mannitol recrystallises from Et<sub>2</sub>O in long fine needles with  $\lambda_{max}$  at 256nm ( $\epsilon$  435) in 95% EtOH, and R<sub>F</sub> 0.21 (1:1 CCl<sub>4</sub>/EtOAc) on TLC Silica Gel G. [Sinclair *Carbohydr Res* 12 150 1970, ORD, CD, NMR, IR, MS: Brecknell et al. *Aust J Chem* 29 1749 1976, *Beilstein* 19/11 V 640.]

**Digitonin** [11024-24-1] **M 1229.3, m >270°(dec)**,  $[\alpha]_{546}^{20} -63^\circ$  (c 3, MeOH). This digitoxin hexa-glycoside can be recrystallised from aqueous 85% EtOH or MeOH/diethyl ether. It is purified by preparative paper chromatography and developed with the upper phase of a mixture of nBuOH/H<sub>2</sub>O/AcOH (4:5:1), and the spot (R<sub>F</sub> 0.36) is eluted with 25% CCl<sub>3</sub>CO<sub>2</sub>H in CHCl<sub>3</sub>. It has also been purified by countercurrent distribution. It forms an ethanolate, and complexes with cholesterol and other sterols. [Ruhstroth-Bauer & Breitenfeld *Hoppe Seyler's Z Physiol Chem* 302 111 1955, Grisvold *J Am Pharm Assoc* 23 664 1934, *Beilstein* 19 IV 1243.]

**Digitoxin** [71-63-6] **M 764.9, m 256-257° (anhydrous)**,  $[\alpha]_D^{20} +16.7^\circ$  (c 1, CHCl<sub>3</sub>),  $+4.8^\circ$  (c 1, dioxane). Digitoxin crystallises from MeOH, aqueous EtOH with 0.5 to 1 H<sub>2</sub>O and from H<sub>2</sub>O as the *dihydrate*. It also crystallises from CHCl<sub>3</sub>/Et<sub>2</sub>O as *anhydrous* crystals. The hydrate dehydrates at 120°/vacuum. Its solubility is 2.5% in CHCl<sub>3</sub>, 1.7% in EtOH, 0.25% in EtOAc, and 0.001% in H<sub>2</sub>O; and has E<sub>1cm</sub><sup>1%</sup> 202.5 at 219-220nm (50% EtOH). [Stoll et al. *Helv Chim Acta* 37 1134 1954, Demoen & Janssen *J Am Pharm Assoc* 42 635 1953, *Beilstein* 18 IV 1478.]

**D(+)-Digitoxose (2,6-dideoxy-D-ribo-hexose)** [527-52-6] **M 148.2, m 110°**, 112°,  $[\alpha]_{546}^{20} +57^\circ$  (c 1, H<sub>2</sub>O). Crystallise D(+)-digitoxose from MeOH/Et<sub>2</sub>O, Et<sub>2</sub>O, EtOAc, EtOAc/Et<sub>2</sub>O/petroleum ether or Me<sub>2</sub>CO/Et<sub>2</sub>O and dry it over P<sub>2</sub>O<sub>5</sub>/vacuum. It has  $[\alpha]_{D}^{20} +45.2^\circ$  (6 minutes) mutarotating to +50.2° (16 hours constant) (c 1.65, H<sub>2</sub>O). [Gut & Prins *Helv Chim Acta* **20** 1229 1947, Bollinger & Ulrich *Helv Chim Acta* **35** 93 1952, NMR of derivs: Tsukamoto et al. *J Chem Soc, Perkin Trans 1* 2621 1988, *Beilstein* **1** IV 4191.]

**Digoxin** [5 $\beta$ ,20(22)-cardenolide-3 $\beta$ ,12 $\beta$ ,14 $\beta$ -triol-3-(O-2,6-dideoxy- $\beta$ -D-ribo-hexopyranosyl-(1 $\rightarrow$ 4)-O-2,6-dideoxy- $\beta$ -D-ribo-hexopyranosyl-(1 $\rightarrow$ 4)-O-2,6-dideoxy- $\beta$ -D-ribo-hexopyranosyl)-oxy-), **12 $\beta$ -isomer**] [20830-75-5] **M 781.0, m 265°(dec)**,  $[\alpha]_{546}^{20} +14.0^\circ$  (c 10, pyridine),  $[\alpha]_{546}^{25} +15.6^\circ$  (c 0.5, CHCl<sub>3</sub>/MeOH 1:1). Crystallise digoxin from aqueous EtOH, aqueous pyridine, EtOH/CHCl<sub>3</sub>, and dry it in a vacuum at 100°. The melting point depends on heating rate, but when placed in a bath at 260° and heated slowly it decomposes at 265°. In EtOH it has  $\lambda_{\max}$  at 220nm ( $\epsilon$  12,800). [Smith *J Chem Soc* 508 1930, X-ray: Go et al. *Cryst Struct Commun* **8** 149, 1031 1979, *Beilstein* **18/4** V 381.] **HIGHLY TOXIC.**

**(-)-2,3:4,6-Di-O-isopropylidene-2-keto-L-gulonic acid monohydrate (-DAG)** [18467-77-1] **M 292.3, m 100-101°**, 103°,  $[\alpha]_{D}^{25} -21.6^\circ$  (c 2.3, MeOH). Dissolve -DAG in Et<sub>2</sub>O, filter. dry (MgSO<sub>4</sub>) it, filter it again and evaporate to give a yellow oil. Addition of one drop of H<sub>2</sub>O induces crystallisation to the *monohydrate*, which also forms rhombic crystals on recrystallisation from 95% EtOH/H<sub>2</sub>O at room temperature. [Flatt et al. *Synthesis* 815 1979, Reichstein & Grussner *Helv Chim Acta* **17** 311 1934, Takagi & Jeffrey *Acta Crystallogr Sect B* **34** 2932 1978, cf *Org Synth* **55** 80 1976, *Beilstein* **19/12** V 520.]

**1,2:5,6-Di-O-isopropylidene-D-mannitol** [1707-77-3] **M 262.3, m 121-125°**, 122°,  $[\alpha]_{D}^{25} +1.2^\circ$  (c 3, H<sub>2</sub>O). Although quite soluble in H<sub>2</sub>O, it gives a purer product when crystallised from this solvent, forming needles [Baer *J Am Chem Soc* **67** 338 1945, NMR: Curtis et al. *J Chem Soc, Perkin Trans 1* 1756 1977]. [*Beilstein* **19/11** V 589.]

**Dulcitol (galactitol)** [608-66-2] **M 182.2, m 188-189°**, b 276-280°/1.1mm, **pK<sup>18</sup> 13.5**. Crystallise dulcitol from water by addition of EtOH. It is optically inactive and is prepared by reduction of D-galactose. Its *hexaacetate* crystallises from EtOH and has **m** 168-169°. [IR: Thompson et al. *Discuss Farad Soc* **9** 222 1950, Wolfrom & Thompson *Methods in Carbohydrate Chemistry* **II** 671963, Academic Press, *Beilstein* **1** IV 2844.]

**meso-Erythritol** [149-32-6] **M 122.1, m 122°**, b 329-331°, **pK<sup>18</sup> 13.9**. *meso*-Erythritol crystallises from distilled water or absolute EtOH and is dried at 60° in a vacuum oven. It sublimes at 110° in a high vacuum. It is optically inactive. [Jeans & Hudson *J Org Chem* **20** 1565 1955, IR: Kuhn *Anal Chem* **22** 276 1950, *Beilstein* **1** IV 2807.]

**Erythryl tetranitrate (Cordite)** [7297-25-8] **M 302.1, m 61°**. Crystallise cordite from EtOH. It explodes on percussion at ~220-460° and is a vasodilator. [*Beilstein* **1** III 2358, **1** IV 2809.]

**D(-)-Fructose** [57-48-7] **M 180.2, m 103-106°**,  $[\alpha]_{546}^{20} -190^\circ$  (after 1 hour, c 10, H<sub>2</sub>O), **pK<sup>25</sup> 12.03**. Dissolve D(-)-fructose in an equal weight of water (charcoal, previously washed with water to remove any soluble material), filter and evaporate under reduced pressure at 45-50° to give a syrup containing 90% of fructose. After cooling to 40°, the syrup is seeded and kept at this temperature for 20-30 hours with occasional stirring. The crystals are removed by centrifugation, washed with a small quantity of water and dried to constant weight under a vacuum over conc H<sub>2</sub>SO<sub>4</sub>. For higher purity, this material is recrystallised from 50% aqueous ethanol. [Tsunami et al. *J Am Chem Soc* **72** 1071 1950]. [*Beilstein* **31** H 321, **1** IV 4401.]

**D(+)-Fucose (6-deoxy-d-galactose)** [3615-37-0] **M 164.2, m 144°**,  $[\alpha]_{546}^{20} +89^\circ$  (after 24 hours, c 10 in H<sub>2</sub>O),  **$\alpha$ -form mutarotates:**  $[\alpha]_{D}^{20} +124^\circ$  (10 minutes) to +75.6° (24 hours) (c 9, EtOH). Crystallise D(+)-fucose from EtOH or 95% EtOH. Its *1,2:3,4-diisopropylidene* derivative has **b** 83-84°/0.45mm and crystallises on seeding with **m** 37° and  $[\alpha]_{D}^{20} -52^\circ$  (melt). [Schmidt *Methods in Carbohydrate Chemistry* **I** 191 1962, Academic Press, Haskett et al. *J Am Chem Soc* **61** 1658 1939, *Beilstein* **31** H 76, **1** IV 4265.]

**D-Galactal (1,5-anhydro-2-deoxy-D-lyxo-hex-1-enitol)** [21193-75-9] **M 146.2, m 100°, 100-102°, 104°, 103-106°,  $[\alpha]_D^{20}$  -21.3° (c 1, MeOH).** Recrystallise D-galactal from EtOAc, EtOH or EtOAc/MeOH. [Overend et al. *J Chem Soc* 675 1950, Wood & Fletcher *J Am Chem Soc* 79 3234 1957, Distler & Jourdan *J Biol Chem* 248 6772 1973, *Beilstein* 17 III/IV 2332.]

**Galactaric acid (mucic acid)** [526-99-6] **M 210.1, m 212-213°(dec)  $pK_{1}^{25}$  3.09 (3.29),  $pK_{2}^{25}$  3.63 (4.41).** Dissolve mucic acid (40g) in the minimum (calculated) volume of N aqueous NaOH (~335ml) without heating, decolorise this with charcoal, filter and precipitate it by adding 5N HCl (~57ml). Cool for 1 hour at 0°, filter off, wash with cold H<sub>2</sub>O and dry it *in vacuo*. All temperatures should be kept below 25°. It is optically inactive. [Barker et al. *J Chem Soc* 4128 1958, Lewis et al. *Methods in Carbohydrate Chemistry* II 39 1963, Academic Press, *Beilstein* 3 IV 1292.]

**D-Galactonic acid** [576-36-3] **M 196.2, m 141° (hydrate), 148° (anhydrous)  $pK_{Est}$  ~3.5.** Crystallise D-galactonic acid from EtOH or aqueous EtOH. It cyclises to *D-galactono-1,4-lactone*, **m 134-136°**, and mutarotates in 1 hour to  $[\alpha]_{346}^{30}$  -92° (c 5, H<sub>2</sub>O). It can also be obtained from the Na salt by adding 10 times its weight of acetic acid, warming till just brown, cooling, filtering off the crystals and drying them. It has **m 145-146°,  $[\alpha]_D^{25}$  -13.6° (c 1, H<sub>2</sub>O, 2 minutes and mutarotates to -57.6°).** [Blackburn & Upton *J Am Chem Soc* 55 2514 1933, *Beilstein* 3 IV 1257, and for the  $\gamma$ -lactone see below.]

**D(-)-Galactono-1,4-lactone** [2782-07-2] **M 178.1, m 134-135°, 134-137°,  $[\alpha]_D^{20}$  -78° (c 5, H<sub>2</sub>O, 1 hour).** Crystallise the lactone from EtOH, aqueous EtOH, MeOH or EtOAc. It is also purified by passage through a column of Amberlite IR-120 (H<sup>+</sup> form), and the effluent and washings are then freeze-dried [Wolfrom & Thompson *Methods in Carbohydrate Chemistry* I 67 1963, Academic Press]. The *5,6-O-isopropylidene-D-galactono-1,4-lactone* is purified by chromatography (EtOAc) and has **m 165-167° (from EtOH/hexane), and 167.5-168.5° (from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>)** [NMR: Copel & Stick *Aust J Chem* 31 1371 1978.] [*Beilstein* 18 III/IV 3026, 18/5 V 18.]

**D(+)-Galactosamine hydrochloride** [1772-03-8] **M 215.6, m 181-185°,  $[\alpha]_D^{25}$  +96.4° (after 24 hours, c 3.2 in H<sub>2</sub>O),  $pK_{Est}$  ~7.7 (free base).** Dissolve the hydrochloride in a small volume of H<sub>2</sub>O. Then add three volumes of EtOH, followed by acetone until faintly turbid and keep overnight in a refrigerator. [Roseman & Ludoweig *J Am Chem Soc* 76 301 1954, *Beilstein* 4 IV 2024.]

**$\alpha$ -D-Galactose** [59-23-4, 3646-73-9 pyranose] **M 180.2, m 167-168°,  $[\alpha]_D^{20}$  +80.4° (after 24 hours, c 4 in H<sub>2</sub>O),  $pK^{25}$  12.48.**  $\alpha$ -D-Galactose is crystallised twice from aqueous 80% EtOH at -10°, then dried in a vacuum oven at 90° over P<sub>2</sub>O<sub>5</sub> for 10 hours. [Link *Biochemical Preparations* 3 75 1953, Hansen et al. *Biochemical Preparations* 4 2 1955.] Also purify it by recrystallising the dried solid (150g) in hot H<sub>2</sub>O (150ml), then adding hot MeOH (250ml) and hot EtOH (500ml), stirring to mix, filtering through a bed of charcoal, and the clear filtrate is stirred to initiate crystallisation. After standing overnight at 10°, the crystals of the  $\alpha$ -anomer are filtered off by suction, washed with MeOH, then EtOH, and dried (yield 130g), and more can be obtained by evaporation of the filtrate and washing as before. [Wolfrom & Thompson *Methods in Carbohydrate Chemistry* I 120 1962, Academic Press, *Beilstein* 1 IV 4336.]

**$\beta$ -D-Galactose** [7296-64-2 (pyranose)] **M 180.2, m 167°,  $[\alpha]_D^{20}$  +52° (initial, c 4 in H<sub>2</sub>O).**  $\alpha$ -D-Galactose (40g) is dissolved in hot H<sub>2</sub>O to establish the equilibrium of  $\alpha$ - and  $\beta$ - anomers; then the solution is cooled to 0° and poured into absolute EtOH (500ml). Stir vigorously and crystallisation occurs within a few minutes, and more rapidly if seeded, filter the crystals immediately (7g,  $[\alpha]_D^{20}$  +65° initial, c 4 in H<sub>2</sub>O). This mixture of  $\alpha$ - and  $\beta$ - anomers is further separated by dissolving in an equal weight of cold H<sub>2</sub>O, filtering and adding to ice cold absolute EtOH (250ml) and stirring for 1 minute when crystals separate, then filter them off. After two such crystallisations, the initial  $[\alpha]_D^{20}$  is +53°. This can be further purified by shaking with 80% EtOH for 2 minutes, filtering, washing with EtOH and Et<sub>2</sub>O, and drying in a vacuum desiccator to give  $\beta$ -D-galactose (15g) with **m 167°,  $[\alpha]_D^{20}$  +52° (initial, c 4 in H<sub>2</sub>O) mutarotating to +80.4°.** Acetylation of D-galactose with hot NaOAc/Ac<sub>2</sub>O gives  **$\beta$ -D-galactopyranoside pentaacetate m 142°,  $[\alpha]_D^{25}$  +25 (c 4 in CHCl<sub>3</sub>).** [Wolfrom & Thompson *Methods in Carbohydrate Chemistry* I 120 1962, Academic Press, *Beilstein* 1 IV 4336.]

**D(+)-Galacturonic acid** [685-73-4, 91510-62-2 (H<sub>2</sub>O)] **M 194.1, m 159-161°**,  $[\alpha]_{\text{D}}^{20} +36^{\circ}$  (c 6, H<sub>2</sub>O, 2h), **pK<sub>Est</sub> ~4.8**. Crystallisation of the acid from 95% EtOH and drying it in a vacuum desiccator (12mm) over P<sub>2</sub>O<sub>5</sub> gives the *monohydrate* mixture of  $\alpha$ - and  $\beta$ - anomers (**mostly  $\alpha$ -**) as white micro needles, which sinter at ~100-111° and melt at 159-160°,  $[\alpha]_{\text{D}}^{20} +107^{\circ}$  (initial, c 4 in H<sub>2</sub>O mutarotating to +51°). [Link & Sell *Biochemical Preparations* 3 74, 78 1953, *Beilstein* 3 IV 2000.] The  **$\beta$ -anomer** is obtained by warming the  $\alpha$ -anomer in EtOH, AcOH or EtOAc and has **m 160° (165°, sinters at 140°)**,  $[\alpha]_{\text{D}}^{20} +27^{\circ}$  (initial, c 2 in H<sub>2</sub>O mutarotating to +55.3° in 24 hours). The *sodium salt* [14984-39-5] **M 216.1** has  $[\alpha]_{\text{D}}^{20} +27^{\circ}$  (c 10 in H<sub>2</sub>O after 5 hours). The *phenylhydrazone* has **m 141°** (from MeOH). [Ehrlich & Schubert *Chem Ber* 62 1974, 2014 1929, Anderson & King *J Chem Soc* 5333 1961, *Beilstein* 3 IV 2001.]

**Genistin (4',5,7-trihydroxyisoflavone-7-D-glucoside)** [529-59-9] **M 432.4, m 256°**,  $[\alpha]_{\text{D}}^{22} -28^{\circ}$  (c 0.6, 0.02N NaOH). Genistin is repeatedly crystallised from hot 80% EtOH/water and treated with charcoal (Nuchar) until free from saponin. The presence of saponin is detected by adding crystals to conc H<sub>2</sub>SO<sub>4</sub> when the citron yellow colour changes to red, then purple. Pure genistin does not change colour. Its UV in 85% EtOH has  $\lambda_{\text{max}}$  at 262.5nm. [Walter *J Am Chem Soc* 63 3273 1941, *Beilstein* 18 III/IV 2732.]

**$\alpha$ -Gentiobiose (amygdalose, 6-O- $\alpha$ -D-glucopyranosyl-D-glucopyranose)** [5995-99-5 (*bi-pyranose*)] **M 342.3, m 86° (2MeOH), 189-195°, 195-197° (anhydrous)**,  $[\alpha]_{\text{D}}^{20} +11^{\circ}$  (after 24 hours, c 4, H<sub>2</sub>O). Crystallise  $\alpha$ -gentiobiose from MeOH (retains solvent of crystallisation). It is best purified by conversion to the  *$\alpha$ -octaacetate*, **m 191-192°** (recrystallise from absolute EtOH or CHCl<sub>3</sub> and excess absolute EtOH),  $[\alpha]_{\text{D}}^{20} +51.6^{\circ}$  (c 4.3, CHCl<sub>3</sub>) [Goldstein & Whelan *Methods in Carbohydrate Chemistry* I 313/1962, Academic Press, Hudson & Johnson *J Am Chem Soc* 39 1271 1917] directly or from the  *$\beta$ -octaacetate* (see below) obtained by acetylation with Ac<sub>2</sub>O and ZnCl<sub>2</sub>, and hydrolysed (see below). It mutarotates from  $[\alpha]_{\text{D}}^{22} +16^{\circ}$  (3 minutes) to +9.5° (c 2, H<sub>2</sub>O). [Reynolds & Evans *J Am Chem Soc* 60 2559 1938, NMR of anomers: Bradbury et al. *Carbohydrate Research* 53 1970 1957, *Beilstein* 17/7 V 203.]

**$\beta$ -Gentiobiose (6-O- $\beta$ -D-glucopyranosyl-D-glucopyranose)** [5996-00-9 (*bi-pyranose*), 554-91-6 (*one open ring*)] **M 342.3, m 190-195°**,  $[\alpha]_{\text{D}}^{20} +8^{\circ}$  (after 6 hours, c 3, H<sub>2</sub>O).  $\beta$ -Gentiobiose is best purified *via* the *octaacetate* which is recrystallised from MeOH, EtOH or better from methyl cellosolve by heating at 80°. The octaacetate (15g) is hydrolysed by suspending it in dry 0.05N NaOMe in MeOH (180ml) for 1 hour with occasional shaking at room temperature. Dilute this with H<sub>2</sub>O to dissolve suspended matter, pass through Amberlite IR-120 and Duolite A-4 columns and the eluate is evaporated under reduced pressure to a syrup. Residual H<sub>2</sub>O is removed by repeated distillation with absolute EtOH under reduced pressure. The syrup is dissolved in methyl cellosolve (~40ml), filtered, nucleated and placed in an oven at 80°. The crystals are filtered off, washed with absolute EtOH (yield 6.7g, 89%), dried and have **m 187-189°**. Further recrystallisation from methyl cellosolve gives **m 190°**, and mutarotates from  $[\alpha]_{\text{D}}^{28} -1.5^{\circ}$  (initial) to +10.6° (final, c 4, H<sub>2</sub>O). The  *$\beta$ -octaacetate* has **m 193°** (crystallised from 95% EtOH) and has  $[\alpha]_{\text{D}}^{20} -5^{\circ}$  (c 1.8, CHCl<sub>3</sub>). [Goldstein & Whelan *Methods in Carbohydrate Chemistry* I 313 1962, Academic Press, *Beilstein* 17/7 V 203.]

**Glucamine (glycamine, 1-amino-1-deoxy-D-glucitol)** [488-43-7] **M 181.2, m 127°**,  $[\alpha]_{\text{D}}^{20} -8^{\circ}$  (c 10, H<sub>2</sub>O), **pK<sub>Est</sub> ~9.0**. Crystallise glucamine from MeOH or aqueous MeOH and store it in a CO<sub>2</sub>-free atmosphere. For the *N-methylglucamine* derivative see below. [Holly et al. *J Am Chem Soc* 72 5461 1950, Karrer et al. *Helv Chim Acta* 20 83 1937.]

**D-Gluconamide** [3118-85-2] **M 197.2, m 142-143°, 144°**,  $[\alpha]_{\text{D}}^{25} +31^{\circ}$  (c 2, H<sub>2</sub>O). Crystallise D-gluconamide from EtOH or 1,2-dimethoxyethane. It mutarotates slowly and hydrolytically in aqueous solution from +30.7° to +13.8° in 80 hours to give *ammonium D-gluconate* which recrystallises from 95% EtOH with **m 153-154°** and  $[\alpha]_{\text{D}}^{25} +11.8^{\circ}$  (c 3, H<sub>2</sub>O). [Wolfrom et al. *J Am Chem Soc* 80 944 1958, *Beilstein* 3 VI 1259.]

**D-Glucono- $\delta$ -lactone** [90-80-2] **M 178.1, m 152-153°**,  $[\alpha]_{\text{D}}^{20} +76^{\circ}$  (c 4, H<sub>2</sub>O). Crystallise D-glucono- $\delta$ -lactone from ethylene glycol monomethyl ether and dry for 1 hour at 110°. It can be freed from other sugars *via*

a column of Celite and charcoal (750g of each, 90 x 7.5cm) which is washed with 0.01N formic acid until the pH of the wash is equal to that of the entering acid. The lactone is applied in H<sub>2</sub>O and eluted with 0.01N formic acid (7L), then eluted with 7.5% EtOH/0.01N formic acid (8L), then 15% EtOH/0.01N formic acid (8L) which removes pentose and isomaltose (the optical rotation of the eluates are used for sugar detection) and finally elution with aqueous formic acid provides glucolactone which is obtained by evaporating or freeze drying. Its solubility in H<sub>2</sub>O is 60% and 1% in EtOH. A solution in H<sub>2</sub>O is slightly acidic, and the lactone dissolves in an equivalent of aqueous NaOH to form **sodium D-gluconate** [527-07-1] **M 218.1**, has **m 200-206°(dec)**,  $[\alpha]_{\text{D}}^{25} +12^{\circ}$  (c 10, H<sub>2</sub>O), **pK<sup>25</sup> 3.6**. [cf p 553, Smith & Whelan *Biochemical Preparations* 10 127 1963, Beilstein 3 IV 1255.]

**α- and β- Glucosamine (2-amino-2-deoxy-D-glucose)** [3416-24-8] **M 179.2**, **m 110°(dec)**,  $[\alpha]_{\text{D}}^{20} +28^{\circ}$  to  $+48^{\circ}$  (c 5, H<sub>2</sub>O), **pK<sup>24</sup> 7.71**. Crystallise the amines from MeOH. The free base has been obtained from the hydrochloride (21.5g, see below) in a mixture of Et<sub>3</sub>N (15ml) and EtOH (125ml) by shaking for 2 days at room temperature, and the solid Et<sub>3</sub>N.HCl is filtered off and the process repeated with more Et<sub>3</sub>N (3-4 times) until the **α-D-glucosamine** (15g) is free from Cl ions. It has **m 88°**,  $[\alpha]_{\text{D}}^{20} +100^{\circ}$  mutarotating to  $+47.5^{\circ}$  (c 1, H<sub>2</sub>O). When Et<sub>2</sub>NH is used as base, the α- to β- conversion is complete giving **β-D-glucosamine**. The *pentaacetate* is purified by dissolving in CHCl<sub>3</sub>, treating with charcoal, drying (MgSO<sub>4</sub>), evaporating the solvent, and adding a little dry Et<sub>2</sub>O to induce crystallisation. It has **m 124-126°**,  $[\alpha]_{\text{D}}^{20} +113^{\circ}$  (c 1, CHCl<sub>3</sub>) after 16 hours in a desiccator. [Leaback *Biochemical Preparations* 10 118 1963.] The *N-acetyl* derivative, **m 203-205°** from MeOH/Et<sub>2</sub>O (dry in vacuum P<sub>2</sub>O<sub>5</sub>) has  $[\alpha]_{\text{D}}^{20} +75^{\circ}$  to  $+41^{\circ}$  (c 2, H<sub>2</sub>O); this derivative can also be purified by dissolving in the minimum volume of H<sub>2</sub>O to which is added 7-8 volumes of EtOH followed by Et<sub>2</sub>O until turbid and keeping at ~20° to crystallise. Wash the crystals with MeOH then Et<sub>2</sub>O and dry *in vacuo* over P<sub>2</sub>O<sub>5</sub>. [Horton *Biochemical Preparations* 11 1 1966.]

**D-Glucosamine hydrochloride** [66-84-2] **M 215.6**, **m >300°**,  $[\alpha]_{\text{D}}^{25} +71.8^{\circ}$  (after 20 hours, c 4, H<sub>2</sub>O), **pK<sup>24</sup> 7.71**. Crystallise the hydrochloride from 3M HCl, water, and finally water/EtOH/acetone as for galactosamine hydrochloride. [Purchase & Braun *Org Synth* 26 36 1946, Stacey & Webber *Methods in Carbohydrate Chemistry* I 228 1962, Academic Press.] The salt has also been purified by dissolving in the minimum volume of boiling H<sub>2</sub>O (charcoal), filtering and adding a large excess of 95% EtOH (~4 volumes) and stirring vigorously for several hours. Collect the crystals after 4-6 hours to give **α-anomer** which mutarotates from  $[\alpha]_{\text{D}}^{25} +100^{\circ}$  to  $+72^{\circ}$  (equilibrium, c 1, H<sub>2</sub>O). A large amount of the **β-anomer** stays in solution. This can be precipitated from the filtrate by adding excess Et<sub>2</sub>O. The mixture of α- plus β-anomers has  $[\alpha]_{\text{D}}^{25.5} +68.8^{\circ}$  (c 4.75, H<sub>2</sub>O, mutarotating to  $+70.1^{\circ}$ ) [Leaback *Biochemical Preparations* 10 118 1963]. *Note that if Et<sub>2</sub>NH is used instead of Et<sub>3</sub>N, conversion to the β-anomer can be complete* (see above). [Stacey et al. *Methods in Carbohydrate Chemistry* I 306 1962, Academic Press, Beilstein 4 IV 2018.]

**α-D-Glucose** [492-62-6] **M 180.2**, **m 83° (monohydrate), 146° (anhydrous), mutarotates from  $[\alpha]_{\text{D}}^{20} +112^{\circ}$  to  $+52.5^{\circ}$  (after 24 hours, c 4, H<sub>2</sub>O), pK<sup>25</sup> 12.46**. Recrystallise α-D-glucose slowly from aqueous 80% EtOH, then dry it over P<sub>2</sub>O<sub>5</sub> *in vacuo*. *Alternatively*, crystallise it from water at 55°, then dry it for 6 hours in a vacuum oven between 60-70°/2mm. Its solubilities are: H<sub>2</sub>O (~50%), EtOH (1.7%). [Hendricks et al. *J Am Chem Soc* 56 99 1934, Beilstein 1 IV 4302.] [For equilibrium forms see Angyal *Adv Carbohydr Chem* 42 15 1984, Angyal & Pickles *Aust J Chem* 25 1711 1972.]

**β-D-Glucose** [50-99-7] **M 180.2**, **m 148-150°**, mutarotates from  $[\alpha]_{\text{D}}^{20} +18.7^{\circ}$  to  $+52.5^{\circ}$  (after 24 hours, c 4, H<sub>2</sub>O). Crystallise β-D-glucose from hot glacial acetic acid or pyridine. Traces of solvent are removed by drying in a vacuum oven at 75° for >3 hours. [Gottfried *Adv Carbohydr Chem* 5 127 1950, Kjaer & Lindberg *Acta Chem Scand* 13 1713 1959, Whistler & Miller *Methods in Carbohydrate Chemistry* I 130 1962, Academic Press, Beilstein 1 IV 4306.] [For equilibrium forms see Angyal *Adv Carbohydr Chem* 42 15 1984, Angyal & Pickles *Aust J Chem* 25 1711 1972.]

**α-D-Glucose pentaacetate** [604-68-2] **M 390.4**, **m 110-111°, 112-113°**,  $[\alpha]_{546}^{20} +119^{\circ}$ ,  $[\alpha]_{\text{D}}^{20} +102^{\circ}$  (c 5, CHCl<sub>3</sub>). Crystallise it from MeOH, EtOH or three recrystallisations from 95% EtOH. [Wolfrom & Thompson *Methods in Carbohydrate Chemistry* II 212 1963, Academic Press, Beilstein 17/7 V 318.]

**$\beta$ -D-Glucose pentaacetate** [604-69-3] **M 390.4, m 131-132<sup>o</sup>,  $[\alpha]_{546}^{20} +5^{\circ}$  (c 5, CHCl<sub>3</sub>).** Crystallise the pentaacetate from MeOH or EtOH. It is best purified by recrystallising 160g from 1L of hot 95% EtOH (charcoal) and filtering hot. It is important that as soon as the temperature of the filtrate cools to  $\sim 20^{\circ}$  it is filtered off. Note that some  $\alpha$ -D-isomer will crystallise out if a prolonged crystallisation period is allowed. Further crystallisation in this manner and drying in a vacuum over CaCl<sub>2</sub> will give pure  $\beta$ -D-anomer which has **m 132<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +4^{\circ}$  (c 5, CHCl<sub>3</sub>).** [Wolfrom & Thompson *Methods in Carbohydrate Chemistry* **II** 212 1963, Academic Press, Krahl & Cori *Biochemical Preparations* **1** 33 1955, *Beilstein* **17/7** V 319.]

**D-Glucose phenylhydrazone** [3713-25-5] **M 358.4, m three forms.** Crystallise the hydrazone from 70% aqueous EtOH or EtOH/Et<sub>2</sub>O. Three forms have been described: “ $\alpha$ ” form **m 159<sup>o</sup>, 160<sup>o</sup>** which mutarotates from  $[\alpha]_{\text{D}}^{20} -87^{\circ}$  to  $-50^{\circ}$  (H<sub>2</sub>O) [Fischer *Ber* **20** 821 1887, Behrends *Justus Liebigs Ann Chem* **353** 106 1907], “ $\beta$ ” form **m 140-141<sup>o</sup>, 142<sup>o</sup>** which mutarotates from  $[\alpha]_{\text{D}}^{20} -2^{\circ}$  to  $-50^{\circ}$  (H<sub>2</sub>O) [Behrends & Lohr *Justus Liebigs Ann Chem* **362** 78 1908], and Skraup’s form **m 115-116<sup>o</sup>** which mutarotates from  $[\alpha]_{\text{D}}^{20} -70^{\circ}$  to  $-47^{\circ}$  (H<sub>2</sub>O) [Skraup *Monatsh Chem* **10** 406 1889, Butler & Crechter *J Am Chem Soc* **51** 3161 1921]. These mutarotate to the **formazan**. [*Beilstein* **1** IV 4322, Mester & Major *J Am Chem Soc* **77** 4297 1955, Stanek et al. *The Monosaccharides*, Academic Press 1963, pp 539-541, 543.]

**D-Glucuronic acid** [6556-12-3] **M 194.1, m 159-161<sup>o</sup>, 165<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +36^{\circ}$  (c 6, H<sub>2</sub>O, 2 hours, mutarotating from +11.5<sup>o</sup>),  $\text{pK}_{\text{a}}^{20}$  3.18.** Crystallise the acid from EtOH or EtOAc, wash it with MeOH and dry it *in vacuo* to give the “ $\beta$ ” form. Heating converts it to the lactone (see below). The *sodium salt monohydrate* [207300-70-7] **M 234.1 has m  $\sim 136$ -138<sup>o</sup>(dec)  $[\alpha]_{\text{D}}^{20} +21^{\circ}$  (c 2, H<sub>2</sub>O after 2hours).** [Sutter & Reichstein *Helv Chim Acta* **21** 1210 1938, *Beilstein* **3** H 886, **3** IV 1997.]

**D-Glucurono-6,3-( $\delta$ )-lactone** [32449-92-6] **M 176.1, m 175-177<sup>o</sup>,  $[\alpha]_{546}^{20} +22^{\circ}$  (after 24 hours, c 10, H<sub>2</sub>O).** Dissolve the lactone or mixture of lactone and acid in H<sub>2</sub>O and concentrate on a steam bath until crystallisation begins. Cool rapidly to room temperature with stirring. After 2 hours the product is filtered off, washed with cold EtOH and dried to **m 174-175<sup>o</sup>** and  $[\alpha]_{\text{D}}^{25} +19.8^{\circ}$  (c 5.2, H<sub>2</sub>O). The amount of free acid can be obtained by titration of an ice-cold aqueous solution with standard alkali. It can be recrystallised from EtOH, EtOH/H<sub>2</sub>O or MeOH, and the highest recorded **m** is 180<sup>o</sup>. [Stacey *J Chem Soc* 1529 1939, Mehlretter et al. *J Am Chem Soc* **73** 2424 1951, *Beilstein* **18** III/IV 3055, **18/5** V 33.]

**D(+)-Glycogen** [9005-79-2] **M 25,000-100,000, m 270-280<sup>o</sup>(dec),  $[\alpha]_{546}^{20} +216^{\circ}$  (c 5, H<sub>2</sub>O).** A 5% aqueous solution (charcoal) of D(+)-glycogen is filtered, and an equal volume of EtOH is added. After standing overnight at 3<sup>o</sup> the precipitate is collected by centrifugation, washed with absolute EtOH, then EtOH/diethyl ether (1:1) and diethyl ether, and dried. [Sutherland & Wosilait *J Biol Chem* **218** 459 1956.]

**Glycyrrhizic acid ammonium salt (3H<sub>2</sub>O)** [53956-04-0] **M 823.0, m 210<sup>o</sup>(dec), 220<sup>o</sup>(dec, sintering at 170<sup>o</sup>),  $[\alpha]_{546}^{20} +60^{\circ}$  (c 1, 50% aqueous EtOH),  $\text{pK}_{\text{Est}} \sim 4.0$ .** Crystallise the ammonium salt from glacial acetic acid, then dissolve it in ethanolic ammonia and evaporate. The *pentahydrate* forms needles from 75% aqueous EtOH, **m 212-217<sup>o</sup>**. The free acid crystallises from glacial acetic acid. [Karrer et al. *Helv Chim Acta* **4** 100 1921, Lithgoe & Tripett *J Chem Soc* 1983, 1987 1950, *Beilstein* **18** IV 5156.]

**Heparin (from pig intestinal mucosa)** [9005-49-6] **M<sub>r</sub>  $\sim 3,000$ , amorphous,  $[\alpha]_{\text{D}}^{20} \sim +55^{\circ}$  (H<sub>2</sub>O).** Most likely contaminants are mucopolysaccharides including heparin sulfate and dermatan sulfate. Purify heparin by precipitation with cetylpyridinium chloride from saturated solutions of high ionic strength. [Cifonelli & Roden *Biochemical Preparations* **12** 12 1968, Wolfrom et al. *J Org Chem* **29** 540 1946, Huggard *Adv Carbohydr Chem* **10** 336-368 1955.]

**Heparin (sodium salt)** [9041-08-1] **M<sub>r</sub>  $\sim 3000$  (low Mol Wt, Bovine), amorphous,  $[\alpha]_{\text{D}}^{20} +47^{\circ}$  (c 1.5, H<sub>2</sub>O).** Dissolve the salt in 0.1M NaCl (1g/100ml) and precipitate it by adding EtOH (150ml). [Wolfrom et al. *J Org Chem* **29** 540 1946, Huggard *Adv Carbohydr Chem* **10** 336-368 1955.]

**Heptyl-β-D-glucopyranoside** [78617-12-6] M 278.4, m 74-77°, 76-77°,  $[\alpha]_D^{20}$  -34.2° (c 5, H<sub>2</sub>O). Purify the glucoside by repeated crystallisation from Me<sub>2</sub>CO which is a better solvent than EtOAc. The *acetate* has m 66-68.5° and  $[\alpha]_D^{20}$  -20.5° (c 4, CHCl<sub>3</sub>) [Pigman & Richtmyer *J Am Chem Soc* **64** 369 1942]. [Beilstein **17** IV 2936.]

**Heptyl-β-D-1-thioglucopyranoside** [85618-20-8] M 294.4, m 98-99°. The *tetraacetyl* derivative is purified by silica gel column chromatography and eluted with a \*C<sub>6</sub>H<sub>6</sub>/Me<sub>2</sub>CO gradient (up to 5% of Me<sub>2</sub>CO) and recrystallises from *n*-hexane as colourless needles m 72-74° (Erbing & Lindberg *Acta Chem Scand* **B30** 611 1976 gave m 69-70°). Hydrolysis with an equivalent of base in methanol gives the desired glucoside. This is a non-ionic detergent for reconstituting membrane proteins and has a critical micelle concentration of 30 mM. [Shimamoto et al. *J Biochem (Tokyo)* **97** 1807 1985; Saito & Tsuchiya *Chem Pharm Bull Jpn* **33** 503 1985.]

**D(-)-Isoascorbic acid (araboascorbic acid, 5R-[R-1,2-dihydroxyethyl-3,4-dihydroxy-5H-furan-2-one])** [89-65-6] M 176.1, m 169°, 174°(0.5 H<sub>2</sub>O, dec),  $[\alpha]_D^{25}$  -16.8° (c 2, H<sub>2</sub>O),  $[\alpha]_D^{20}$  +77° (c 2, Me<sub>2</sub>CO, acetonilidene derivative), pK<sup>18</sup> 3.99 (4.23). Crystallise D(-)-isoascorbic acid from H<sub>2</sub>O, EtOH or dioxane.  $\lambda_{\max}$  is at 245nm with  $\epsilon$  7500 (EtOH). [Reichstein et al. *Helv Chim Acta* **17** 510, 516 1934, Heslop et al. *J Chem Soc* 225 1944, Beilstein **18** III/IV 3037, **18/5** V 26.]

**2-Keto-L-gulonic acid (xylo-2-hexulonic acid)** [526-98-7] M 194.1, m 159-162°, 171°,  $[\alpha]_D^{20}$  -34.8° (c 2, MeOH),  $[\alpha]_D^{18}$  -48° (c 1, H<sub>2</sub>O), pK<sup>Est</sup> ~2.2. Crystallise 2-keto-L-gulonic acid from half its weight of water, wash it with Me<sub>2</sub>CO and dry it *in vacuo*. [Reichstein et al. *Helv Chim Acta* **17** 311 1934, NMR: Crawford et al. *J Am Chem Soc* **102** 2220 1980, Beilstein **3** IV 1985.]

**Lactobionic acid (4-O-β-D-galactopyranosyl-D-gulonic acid)** [96-82-2] M 358.3, m 128-130°, mutarotates from  $[\alpha]_D^{20}$  +53° to +22.6° (c 3, after 4 hours in H<sub>2</sub>O), pK<sup>25</sup> 3.6. Crystallise lactobionic acid from water by addition of EtOH. [NMR: Taga et al. *Bull Chem Soc Jpn* **51** 2278 1978, Beilstein **17** III/IV 3392, **17/7** V 436.]

**α-Lactose (H<sub>2</sub>O)** [63-42-3] M 360.3, m 220°(dec), 253-255° (252.4°),  $[\alpha]_D^{20}$  +52.3° (c 4.2, H<sub>2</sub>O), pK<sup>25</sup> 12.2 (OH). α-Lactose crystallises from water below 93.5° as the *hydrate* which can be dried at 80°/14mm. [Horst *Rec Trav Chim, Bas* **72** 878 1953, Beilstein **17** III/IV 3066.]

**Lactulose (4-O-β-D-galactopyranosyl-D-fructose)** [4618-18-2] M 342.2, m 168.5-169°(dec),  $[\alpha]_{D46}^{20}$  -57° (c 1, H<sub>2</sub>O). Crystallise lactulose from MeOH or 50% MeOH. It mutarotates from  $[\alpha]_D^{20}$  -11.9° to -50.7° (c 1, H<sub>2</sub>O). [Montgomery & Hudson *J Am Chem Soc* **32** 2101, 2104 1930, Beilstein **17** III/IV 3094, **17/7** V 214.] NMR in Me<sub>2</sub>SO at 24° shows 0% α-pyranose, 27% β-pyranose, 20% α-furanose and 52% β-furanose forms [Angyal *Adv Carbohydr Chem* **42** 15 1984].

**Lanatoside A (digitoxigenin nonoacetyl-tetraglycoside)** [17575-20-1] M 969.1, m 245-248°,  $[\alpha]_D^{20}$  +32° (EtOH). Crystallise lanatoside A from MeOH or 95% aqueous EtOH. Its solubility is 1/16000 (H<sub>2</sub>O), 1/20 (MeOH), 1/40 (EtOH) and 1/225 (CHCl<sub>3</sub>). [Stoll et al. *Helv Chim Acta* **16** 1049 1933, Kuhn et al. *Helv Chim Acta* **45** 881 1962, Beilstein **18** III/IV 1480.] It is **cardiotonic**.

**Lanatoside B (gitoxigenin nonoacetyl-tetraglycoside)** [17575-21-2] M 985.1, m 233°(dec), 245-248°(dec),  $[\alpha]_D^{20}$  +35° (MeOH). Crystallise lanatoside B from MeOH. Its solubility is: 1/20 (MeOH), 1/40 (EtOH), 1/5600 (CHCl<sub>3</sub>), and is insoluble in H<sub>2</sub>O. [Stoll et al. *Helv Chim Acta* **16** 1049 1933, Miyatake et al. *Chem Pharm Bull Jpn* **5** 171 1957, Beilstein **18** III/IV 2465.] It is **cardiotonic**.

**Lanatoside C (digoxigenin nonoacetyl-tetraglycoside)** [17575-22-3] M 297.1, m 246-248°,  $[\alpha]_D^{20}$  +33.5° (c

**1.85, 95% EtOH).** Crystallise lanatoside C from MeOH. Its solubility is : 1/17000 (H<sub>2</sub>O), 1/45 (EtOH) and 1/1500-2000 (CHCl<sub>3</sub>). [Stoll et al. *Helv Chim Acta* **16** 1049 1933, Stoll & Kreis *Helv Chim Acta* **35** 1318 1952, Okada et al. *Chem Pharm Bull Jpn* **23** 2039 1975, *Beilstein* **18** III/IV 2455.] It is **cardiotonic**.

**$\alpha$ -L(+)-Lyxose** [1949-78-6] **M 150.1, m 106-108<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +14^{\circ}$  after 1 hour (c 6, H<sub>2</sub>O).** The  $\alpha$ -anomer crystallises from propan-1-ol or EtOH, and the  $\beta$ -anomer crystallises from propan-2-ol. The *2,4-dinitrophenylhydrazone* has **m 171-172<sup>o</sup>** and  **$[\alpha]_{\text{D}}^{20} -31^{\circ}$**  (pyridine). In D<sub>2</sub>O it has 21% of the  $\alpha$ -pyranose form. The *2-methyl ether* has **m 120-121<sup>o</sup>** and  **$[\alpha]_{\text{D}}^{20} +6^{\circ}$**  (c 1, H<sub>2</sub>O). [Angyal *Adv Carbohydr Chem* **42** 15 1984, Bently *J Am Chem Soc* **79** 1720 1959, *Beilstein* **1** I 439, **1** IV 4232.]

**$\beta$ -D-Lyxose** [1114-34-7] **M 150.1, m 118-119<sup>o</sup>, 120-122<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} -14^{\circ}$  (c 4, H<sub>2</sub>O),  $\alpha$ -anomer has m 105-107<sup>o</sup> mutarotates from  $[\alpha]_{\text{D}}^{20} +5.6^{\circ}$  to  $-18.8^{\circ}$  (c 4, H<sub>2</sub>O).** Crystallise  $\beta$ -D-lyxose from EtOH or aqueous 80% EtOH by slow crystallisation. Dry it under vacuum at 60<sup>o</sup>, and store it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub> or CaSO<sub>4</sub>. [*Beilstein* **1** IV 4230, Overend et al. *J Chem Soc* 3496 1961.] <sup>1</sup>H NMR in D<sub>2</sub>O has  $\delta$ :  $\alpha$ -H (pyranose) 5.39 (*J* 4.0 Hz),  $\beta$ -H (pyranose) 5.19 (*J* ~0 Hz),  $\alpha$ -H (furanose) 5.26 (*J* 4.0 Hz),  $\beta$ -H (furanose) 5.24 (*J* 4.5 Hz), and at 31<sup>o</sup> in D<sub>2</sub>O it consists of 70%  $\alpha$ -pyranose, 28%  $\beta$ -pyranose, 1.5%  $\alpha$ -furanose and 0.5%  $\beta$ -furanose [Angyal & Pickles *Aust J Chem* **25** 1711 1972].

**Maltose (H<sub>2</sub>O) (4-O- $\alpha$ -D-glucopyranosyl-D-glucose)** [6363-53-7] **M 360.3, m 109-110<sup>o</sup>, 118<sup>o</sup>, mutarotates from  $[\alpha]_{\text{D}}^{20} +111.7^{\circ}$  to  $+130.4^{\circ}$  (c 4, H<sub>2</sub>O).** Purify maltose by chromatography from aqueous solution on to a charcoal/Celite (1:1) column, wash it with water to remove glucose and other monosaccharides, then elute it with aqueous 75% EtOH. Crystallise it from water, aqueous EtOH or EtOH containing 1% nitric acid. Dry it as the *monohydrate* at room temperature under vacuum over H<sub>2</sub>SO<sub>4</sub> or P<sub>2</sub>O<sub>5</sub>. Also purify it by dissolving it in MeOH, evaporating to a syrup which on standing for 12 hours in contact with 1/10<sup>th</sup> its volume of H<sub>2</sub>O gives crystals of the *monohydrate*. Its iodine number is 55.5. The **osazone** has **m 200<sup>o</sup>(dec)** and  **$[\alpha]_{\text{D}}^{20} +58^{\circ}$**  (c 1.4, H<sub>2</sub>O). [Howarth et al. *J Chem Soc* 793 1937, *Beilstein* **17** III/IV 3057, **17** V 189.]

**D-Mannitol** [69-65-8] **M 182.2, m 166.1<sup>o</sup>, 168-170<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +29^{\circ}$  (c 10, after 1 hour in 8% borax solution),  $[\alpha]_{\text{D}}^{22} -4.0^{\circ}$  (c 0.5, DMF), pK<sup>18</sup> 13.5.** D-Mannitol is crystallised from EtOH, MeOH or H<sub>2</sub>O and dried at 100<sup>o</sup>. [Thomson *Acta Chem Scand* **6** 270, 279, 280 1952, *Beilstein* **1** IV 2841.]

**Mannitol hexanitrate** [15825-70-4] **M 452.2, m 112-113<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +45.8^{\circ}$  (c 1.4, EtOH).** The hexanitrate crystallises from EtOH or dilute EtOH as silky white needles. **EXPLOSIVE (on detonation).** [Hayward *J Am Chem Soc* **73** 1974 1951, Patterson & Todd *J Chem Soc* 2876 1929, *Beilstein* **1** IV 2841.]

**$\alpha$ -D(+)-Mannose** [3458-28-4] **M 180.2, m 132<sup>o</sup>,  $[\alpha]_{\text{D}}^{20}$  mutarotates from  $+29.9^{\circ}$  to  $+14^{\circ}$  (c 4, H<sub>2</sub>O).** Crystallise  $\alpha$ -D(+)-mannose repeatedly from EtOH, aqueous 80% EtOH, AcOH or MeOH/propan-2-ol and then dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub> at 60<sup>o</sup>. [For <sup>1</sup>H NMR and equilibrium forms; Angyal *Adv Carbohydr Chem* **42** 15 1984, Angyal & Pickles *Aust J Chem* **25** 1711 1972, *Beilstein* **1** IV 4328.]

**D(+)-Melezitose (H<sub>2</sub>O) (O- $\alpha$ -D-glucopyranosyl-(1  $\rightarrow$  3)- $\beta$ -fructofuranosyl- $\alpha$ -D-glucopyranose** [597-12-6] **M 540.5, m 153-154<sup>o</sup>(dec), 2H<sub>2</sub>O, 160<sup>o</sup>(dec),  $[\alpha]_{\text{D}}^{20} +88^{\circ}$  (c 2, H<sub>2</sub>O for dihydrate) and  $[\alpha]_{\text{D}}^{20} +91.7^{\circ}$  (c 2, H<sub>2</sub>O for anhydrous).** D(+)-Melezitose crystallises from aqueous EtOH as the *monohydrate* and water as the *dihydrate*, and is then dried at 110<sup>o</sup> (anhydrous). It is also purified by dissolving in an equal volume of H<sub>2</sub>O, filtering into a crystallising dish and allowing to stand (loosely covered) for several weeks undisturbed at 20<sup>o</sup>. The crystals of clear prisms are wiped carefully and dried in air. They effloresce at once losing 3.35% of their weight, and after 3 days in air the loss is for 1H<sub>2</sub>O from the dihydrate. Drying at 110<sup>o</sup> for 6 hours *in vacuo* yields the *anhydrous* form which after 2 days in air it absorbs H<sub>2</sub>O to give the *monohydrate*. The *monohydrate* is also obtained by dissolving it in an equal weight of H<sub>2</sub>O at 60<sup>o</sup> and adding 4 volumes of 95% EtOH and drying in air overnight. [Richtmeyer & Hudson *J Org Chem* **11** 610 1946, *Beilstein* **17** III/IV 3815, **17** V 414.]

**D(+)-Melibiose (2H<sub>2</sub>O) (6-O- $\alpha$ -D-galactopyranosyl-D-glucose)** [585-99-9, *monohydrate* 66009-10-7] **M**

**360.3, m 84-85°, 178-181°, 184-185°,  $[\alpha]_D^{20} +135^\circ$  (c 5, after 10 hours H<sub>2</sub>O).** D(+)-Melibiose crystallises as a *hydrate* from water or aqueous EtOH. The  $\alpha$ -**anomer** is obtained by recrystallising 1g from a mixture of 0.35ml of H<sub>2</sub>O and 0.2ml of EtOH. It crystallises easily with **m** 179° and mutarotates from  $[\alpha]_D^{20} +166^\circ$  to  $+142.3^\circ$  (220 minutes, c 4, H<sub>2</sub>O). Crystallisation from MeOH gives the *anhydrous* form which hydrates to the monohydrate in air. The  $\beta$ -**anomer dihydrate** **m** 85-86° mutarotates from  $[\alpha]_D^{20} +123.5^\circ$  to  $+143.1^\circ$  (c 4, H<sub>2</sub>O for anhydrous). [Fletcher & Diehl *J Am Chem Soc* **74** 5774 1952, *Beilstein* **17** III/IV 3075, 17/7 V 206.]

**N-Methyl-D(-)-glucamine (Meglumine)** [6284-40-8] **M 195.2, m 128-129°,  $[\alpha]_{546}^{20} -19.5^\circ$  (c 2, H<sub>2</sub>O), pK<sup>28</sup> 9.62.** Crystallise N-methyl-D(-)-glucamine from MeOH. Its solubility in H<sub>2</sub>O is 10%. [Karrer & Herkenrath *Helv Chim Acta* **20** 83 1957 also for other N-alkyl derivatives, *Beilstein* **4** IV 1914.]

**N-Methyl  $\alpha$ -L-glucosamine** [42852-95-9] **M 193.2, m glass,  $[\alpha]_D^{25} -65^\circ$  (c 1, MeOH) pK<sub>Est</sub> ~9.** The *hydrochloride* crystallises from EtOH as hygroscopic needles with **m** 160-163°,  $[\alpha]_D^{25} +103^\circ$  mutarotating to  $-88^\circ$  after 24 hours (c 0.6, H<sub>2</sub>O), and gives the free base as a glass. The *Pentaacetate* crystallises from CHCl<sub>3</sub>/Et<sub>2</sub>O with **m** 160.5-161.5°,  $[\alpha]_D^{25} -100^\circ$  (c 0.7, CHCl<sub>3</sub>), and the *N-acetate* crystallises from CHCl<sub>3</sub>/MeOH with **m** 165-166°,  $[\alpha]_D^{25} -51^\circ$  (c 0.4, H<sub>2</sub>O). [Kuehl et al. *J Am Chem Soc* **68** 536 1946, **69** 3032 1947, Lemieux & Wolfrom *Adv Carbohydr Chem* **3** 337 1948, *Beilstein* **4** IV 2032.]

**Methyl  $\alpha$ -D-glucoside (methyl  $\alpha$ -D-glucopyranoside)** [97-30-3] **M 194.2, m 168°, 166-169°,  $[\alpha]_D^{20} +158.9^\circ$  (c 10, H<sub>2</sub>O), pK<sup>25</sup> 13.71.** Crystallise methyl  $\alpha$ -D-glucoside from MeOH or EtOH. Its solubility in H<sub>2</sub>O is 10%. [Ferrier et al. *Carbohydr Research* **27** 55, 59 1973, *Beilstein* **17/7** V 13.]

**Methyl  $\beta$ -D-glucoside (methyl  $\beta$ -D-glucopyranoside)** [7000-27-3] **M 203.2 (0.5 H<sub>2</sub>O), m 107-109°,  $[\alpha]_D^{20} -33^\circ$  (c 10, H<sub>2</sub>O).** Crystallise methyl  $\beta$ -D-glucoside from MeOH or EtOH. Its solubility in H<sub>2</sub>O is 10%. [Ferrier et al. *Carbohydr Research* **27** 55, 59 1973, *Beilstein* **17/7** V 10.]

**4-Methylumbellifer-7-yl- $\alpha$ -D-glucopyranoside** [17833-43-1] **M 338.3, m 209-210°, 221-222°,  $[\alpha]_D^{20} +162^\circ$  (c 0.5, pyridine).** Recrystallise 4-methylumbellifer-7-yl- $\alpha$ -D-glucopyranoside from hot H<sub>2</sub>O or EtOH. [Courtin-Duchateau & Veyrière *Carbohydr Research* **65** 23, 29 1978, *Beilstein* **18** IV 443.]

**4-Methylumbellifer-7-yl- $\beta$ -D-glucopyranoside** [18997-57-4] **M 338.3, m 211-213°, 211°,  $[\alpha]_D^{20} -68^\circ$  (c 0.5, pyridine),  $-89.5^\circ$  (c 0.5, H<sub>2</sub>O for half hydrate).** 4-Methylumbellifer-7-yl- $\beta$ -D-glucopyranoside crystallises as the half hydrate from hot H<sub>2</sub>O. [Constantzas & Kocourek *Col Czech Chem Commun* **24** 1099 1959, De Re et al. *Ann Chim (Rome)* **49** 2089 1959, Courtin-Duchateau & Veyrière *Carbohydr Research* **65** 23, 29 1978, *Beilstein* **18** III/IV 5152, **8** IV 433, **18/7** V 616.]

**Naringin (4',5,7-trihydroxyflavanone 7-rhamnoglucoside)** [10236-47-2] **M 580.5, m ~83° (6H<sub>2</sub>O), 171° (2H<sub>2</sub>O),  $[\alpha]_D^{19} -90^\circ$  (c 1, EtOH),  $[\alpha]_{546}^{20} -107^\circ$  (c 1, EtOH).** This bitter principle from grape juice crystallises from water to give the hydrate with 6-8 H<sub>2</sub>O which when dried at 110° gives the *dihydrate*. Its solubility in H<sub>2</sub>O is 0.1% at 40° and 10% at 75°. The *2,4-dinitrophenylhydrazone* crystallises from aqueous dioxane with **m** 246-247° [Douglass et al. *J Am Chem Soc* **73** 4023 1951]. [Pulley & von Loesecke *J Am Chem Soc* **61** 175 1939, *Beilstein* **18** III/IV 2637, **18** V 528.]

**2-Nitrophenyl- $\beta$ -D-galactopyranoside** [369-07-3] **M 301.3, m 185-190°, 193°, 193-194°,  $[\alpha]_D^{18} -51.9^\circ$  (c 1, H<sub>2</sub>O).** Purify 2-nitrophenyl- $\beta$ -D-galactopyranoside by recrystallisation from EtOH. [Seidman & Link *J Am Chem Soc* **72** 4324 1950, Snyder & Link *J Am Chem Soc* **75** 1758 1953]. It is a chromogenic substrate for  $\beta$ -galactosidases [Jagota et al. *J Food Sci* **46** 161 1981]. [*Beilstein* **17/7** V 52.]

**4-Nitrophenyl- $\alpha$ -D-galactopyranoside** [7493-95-0] **M 301.3, m 166-169°, 173°,  $[\alpha]_D^{25} +248^\circ$  (c 1, H<sub>2</sub>O).** Purify 4-nitrophenyl- $\alpha$ -D-galactopyranoside by recrystallisation from H<sub>2</sub>O or aqueous EtOH. The *monohydrate* has **m** 85° which resolidifies and melts again at 151-152° (the *hemihydrate*), then resolidifies again and melts at 173° to give the *anhydrous* form. Drying the monohydrate at 60° yields the hemihydrate, and drying at 100°

gives the anhydrous compound. The *tetraacetate* has *m* 147° after drying at 100°. [Jermyn *Aust J Chem* **15** 569 1962, Helfreich & Jung *Justus Liebigs Ann Chem* **589** 77 1954.] It is a substrate for  $\alpha$ -galactosidase [Dangelmaier & Holmsen *Anal Biochem* **104** 182 1980]. [Beilstein **17/7** V 55.]

**4-Nitrophenyl- $\beta$ -D-galactopyranoside** [3150-24-1] *M* 301.3, *m* 178°, 178-181°, 181-182°,  $[\alpha]_D^{20}$  -83° (c 1, H<sub>2</sub>O). Purify the galactoside by recrystallisation from EtOH. [Horikoshi *J Biochem (Tokyo)* **35** 39 1042, Goebel & Avery *J Exptl Medicine* **50** 521 1929, Snyder & Link *J Am Chem Soc* **75** 1758.] It is a chromogenic substrate for  $\beta$ -galactosidases [Buoncore et al. *J Appl Biochem* **2** 390 1980]. [Beilstein **17/7** V 55.]

**4-Nitrophenyl- $\alpha$ -D-glucopyranoside** [3767-28-0] *M* 301.3, *m* 206-212°, 216-217° (sinters at 210°),  $[\alpha]_D^{20}$  +215° (c 1, H<sub>2</sub>O). Purify 4-nitrophenyl- $\alpha$ -D-glucopyranoside by recrystallisation from H<sub>2</sub>O, MeOH or EtOH. [Jermyn *Aust J Chem* **7** 202 1954, Montgomery et al. *J Am Chem Soc* **64** 690 1942.] It is a chromogenic substrate from  $\alpha$ -glucosidases [Oliviera et al. *Anal Biochem* **113** 188 1981], and is a substrate for glucansucrases [Binder & Robyt *Carbohydr Research* **124** 287 1983]. [Beilstein **17/7** V 53.]

**4-Nitrophenyl- $\beta$ -D-glucopyranoside** [2492-87-7] *M* 301.2, *m* 164°, 164-165°, 165°,  $[\alpha]_D^{20}$  -107° (c 1, H<sub>2</sub>O). Purify 4-nitrophenyl- $\beta$ -D-glucopyranoside by recrystallisation from EtOH or H<sub>2</sub>O. [Montgomery et al. *J Am Chem Soc* **64** 690 1942, Snyder & Link *J Am Chem Soc* **75** 1758 1953.] It is a chromogenic substrate for  $\beta$ -glucosidases [Weber & Fink *J Biol Chem* **255** 9030 1980]. [Beilstein **17/7** V 53.]

***N*-Nonanoyl-*n*-methylglucamine (Mega-9)** [85261-19-4] *M* 335.4, *m* 87-89°. It is a non-ionic detergent which is purified as for *n*-decanoyl-*N*-methylglucamine above. [Hildreth *Biochem J* **207** 363 1982.]

**Nonyl- $\beta$ -D-glucopyranoside** [69984-73-2] *M* 306.4, *m* 67.5-70°, 70-71°,  $[\alpha]_D^{20}$  -34.4° (c 5, H<sub>2</sub>O),  $[\alpha]_D^{25}$  -28.8° (c 1, MeOH). Purify nonyl- $\beta$ -D-glucopyranoside by recrystallisation from Me<sub>2</sub>CO or hexane/Et<sub>2</sub>O and store it in well-stoppered containers as it is *hygroscopic*. [Pigman & Richtmyer *J Am Chem Soc* **64** 369 1942.] It is a UV transparent non-ionic detergent for solubilising membrane proteins [Schwendener et al. *Biochem Biophys Res Commun* **100** 1055 1981]. [Beilstein **17** III/IV 2937, 17/7 V 39.]

**Octyl- $\beta$ -D-glucopyranoside** [29836-26-8] *M* 292.4, *m* 62-65°, 63.8-65°,  $[\alpha]_D^{20}$  s-34° (c 4, H<sub>2</sub>O). Purify octyl- $\beta$ -D-glucopyranoside by recrystallisation from Me<sub>2</sub>CO. It is *hygroscopic* and should be stored in a well-stoppered container. [Noller & Rockwell *J Am Chem Soc* **60** 2076 1938, Pigman & Richtmyer *J Am Chem Soc* **64** 369 1942.] It is a UV transparent non-ionic dialysable detergent for solubilising membrane proteins. The  $\alpha$ -*D*-isomer with  $[\alpha]_D^{20}$  +118° (c 1, MeOH) has similar solubilising properties. [Lazo & Quinn *Anal Biochem* **102** 68 1980, Stubbs et al. *Biochim Biophys Acta* **426** 46 1976, Beilstein **17/7** V 38.]

**Pectic acid** [9046-40-6] *M<sub>r</sub>* (C<sub>6</sub>H<sub>8</sub>O<sub>6</sub>)<sub>n</sub> ~500,000, **amorphous**,  $[\alpha]_D$  +250° (c 1, 0.1M NaOH). Citrus pectic acid (500g) is refluxed for 18 hours with 1.5L of 70% EtOH, and the suspension is filtered hot. The residue is washed with hot 70% EtOH and finally with ether. It is dried in a current of air, ground and dried for 18 hours at 80° under vacuum. [Morell & Link *J Biol Chem* **100** 385 1933.] It can be further purified by dispersing it in water and adding just enough dilute NaOH to dissolve the pectic acid, then passing the solution through columns of cation- and anion-exchange resins [Williams & Johnson *Ind Eng Chem (Anal Ed)* **16** 23 1944], and precipitating with two volumes of 95% EtOH containing 0.01% HCl. The precipitate is worked with 95% EtOH, then Et<sub>2</sub>O, dried and ground. [Rees & Walsh *Angew Chem Int Edn* **16** 214 1977, Rees *Adv Carbohydr Chem* **24** 267 1969.]

**Pectin (1-4 linked)** [9000-69-5] *M<sub>r</sub>* 30,000-100,000, **amorphous**. Dissolve the pectin in hot water to give a 1% solution, then cool, and make it to about 0.05M in HCl by addition of conc HCl, and precipitate it by pouring it slowly, with vigorous stirring into two volumes of 95% EtOH. After standing for several hours, the pectin is filtered through a nylon cloth, then redispersed in 95% EtOH and stood overnight. The precipitate is filtered off, washed with EtOH/Et<sub>2</sub>O, then Et<sub>2</sub>O and dried in air. [Rees & Walsh *Angew Chem, Int Edn* **16** 214 1977, Rees *Adv Carbohydr Chem* **24** 267 1969.]

**Pentaerythritol (2,2-bis[hydroxymethyl]-1,3-propanediol)** [115-77-5] **M 136.2, m 260.5°, 268-269°**. Reflux pentaerythritol with an equal volume of MeOH, then cool, and the precipitate is collected and dried at 90°. It can also be crystallised from dilute aqueous HCl. After sublimation under high vacuum at 200° it has **m 265.5°**. Its solubility in H<sub>2</sub>O is 10%. [Beilstein **18** III 2361, **1** IV 2812.]

**Pentaerythritol tetraacetate** [597-71-7] **M 304.3, m 83-84°, 84-86°**. Crystallise pentaerythritol tetraacetate from hot water, then leach it with cold water until the odour of acetic acid is no longer detectable. It also crystallises from 95% EtOH after dissolving in CHCl<sub>3</sub>, washing with saturated NaHCO<sub>3</sub>, then H<sub>2</sub>O, drying over anhydrous CaCl<sub>2</sub> and evaporating. It has been prepared by acetolysis of the tetranitrate in 95% yield [Wolf from et al. *J Am Chem Soc* **73** 874 1951]. [Breusch & Oguzer *Chem Ber* **88** 1511 1955, LeFèvre et al. *J Chem Soc* **16** 1958, Beilstein **1** IV 1812, **2** IV 264.]

**Pentaerythrityl laurate (pentaerythrityl tetra-*n*-dodecanoate)** [13057-50-6] **M 864.6, m 50°**. Crystallise the laurate from Me<sub>2</sub>CO, Et<sub>2</sub>O or petroleum ether. [Breusch & Oguzer *Chem Ber* **88** 1511 1955.]

**Pentaerythritol tetranitrate** [78-11-5] **M 316.2, m 140.1°**. Crystallise pentaerythritol tetranitrate from acetone or acetone/EtOH. When crystallised from H<sub>2</sub>O at 0°, it may have **m 26-28°** (hydrate?). It detonates more easily than TNT on percussion. The *O*-acetate, when crystallised from EtOH, has **m 87-88°**. Although it has been distilled at 60°/2mm, distillation should **NOT** be attempted as it is **VERY EXPLOSIVE**. It is a vasodilator. [Marans et al. *J Am Chem Soc* **76** 1304 1954, Camp et al. *J Am Chem Soc* **77** 751 1955, Beilstein **1** IV 2816, **2** IV 264.]

**2-Phenylethyl-β-D-thiogalactoside** [63407-54-5] **M 300.4, m 108°, [α]<sub>D</sub><sup>25</sup> -32.2° (c 5, MeOH)**. Recrystallise the thiogalactoside from H<sub>2</sub>O and dry in air to give the 1.5.H<sub>2</sub>O which has **m 80°**. The anhydrous surfactant is obtained by drying it at 78° over P<sub>2</sub>O<sub>5</sub>. [Heilfrich & Türk *Chem Ber* **89** 2215 1856.]

**Phenyl-β-D-galactopyranoside** [2818-58-8] **M 256.3, m 153-154°, 146-148°, 155-156°(dried at 105°), [α]<sub>D</sub><sup>20</sup> -42° (c 1, H<sub>2</sub>O)**. Recrystallisation of phenyl-β-D-galactopyranoside from H<sub>2</sub>O gives the 0.5H<sub>2</sub>O. [Conchie & Hay *Biochem J* **73** 327 1959, IR: Whistler & House *Analyt Chem* **25** 1463 1953.] It is an acceptor substrate for fucosyltransferase [Chester et al. *Eur J Biochem* **69** 583 1976]. [Beilstein **17**/7 V 46.]

**Phenyl-β-D-glucopyranoside** [1464-44-4] **M 256.3, m 174-175°, 174-176°, 176°, 176-178°, [α]<sub>D</sub><sup>20</sup> -72.2° (c 1 for dihydrate, H<sub>2</sub>O)**. Phenyl-β-D-glucopyranoside recrystallises from H<sub>2</sub>O with 2H<sub>2</sub>O and can be dried *in vacuo* at 100°/P<sub>2</sub>O<sub>5</sub>. The dry preparation has [α]<sub>D</sub><sup>25</sup> -70.7° (c 2, H<sub>2</sub>O). [Robertson & Waters *J Chem Soc* 2729 1930, IR: Bunton et al. *J Chem Soc* 4419 1955, Takahashi *Yakugaku Zasshi (J Pharm Soc Jpn)* **74** 7436 1954, Whittler & House *Anal Chem* **25** 1463 1953, UV: Lewis *J Am Chem Soc* **57** 898 1935.] It is a substrate for β-D-glucosidase [deBryne *Eur J Biochem* **102** 257 1979]. [Beilstein **17** III/V 2946, **17**/7 V 46.]

**Phlorizin (2H<sub>2</sub>O) [phloretin 2'-O-β-D-glucoside]** [60-81-1] **M 472.5, m 110°, [α]<sub>D</sub><sup>20</sup> -62° (c 3.2, EtOH)**. Phlorizin crystallises as the *dihydrate* from water and causes glycosuria. [Brazy & Dennis *Am J Physiol* **234** 1279 1978, Zemplén & Bognár *Chem Ber* **17B** 1040 1943, Beilstein **17**/7 V 177.]

**D(+)-Raffinose (5H<sub>2</sub>O) (Melitose, 6-O-α-D-galactopyranosyl-D-glucopyranosyl-β-D-fructo-furanose** [17629-30-0 (5H<sub>2</sub>O), 512-69-6 (anhydrous)] **M 594.5, m 80°, 80-82°, [α]<sub>D</sub><sup>20</sup> +124° (c 10, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> +105° (c 1 for pentahydrate, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 12.40, pK<sub>2</sub><sup>25</sup> 13.44, pK<sub>3</sub><sup>25</sup> 13.52**. D(+)-Raffinose crystallises from H<sub>2</sub>O, 90% aqueous EtOH or MeOH as the *pentahydrate*. The *anhydrous* sugar has **m 132-135°**. It has R<sub>F</sub> 0.8 on TLC (Silica Gel, and 1:3:3 CHCl<sub>3</sub>/butanone:/MeOH). The *undecaacetate* has been purified through an alumina column by elution with CHCl<sub>3</sub>, and recrystallised from EtOH/MeOH/H<sub>2</sub>O (3:2:5), with **m 99-100°, [α]<sub>D</sub><sup>20</sup> +92.8° (c 5.14, EtOH)**. [pK : Coccioli & Vicedomini *Ann Chim (Rome)* **66** 269, 275 1976, <sup>1</sup>H NMR: Suami et al. *Carbohydr Research* **26** 234 1973, Beilstein **17** III/IV 3801, **17**/8 V 403.]

**L(+)-α-Rhamnose (H<sub>2</sub>O) (6-deoxy-L-mannose)** [10030-85-0 (H<sub>2</sub>O), 3615-41-6 (anhydrous)] **M 182.2, m 90-92°, 101°, 105°, [α]<sub>D</sub><sup>18</sup> -6.8° mutarotating to +9.1° (c 1, H<sub>2</sub>O)**. Crystallise the rhamnose from H<sub>2</sub>O or

EtOH. It crystallises easily as the *monohydrate* by evaporating a solution in MeOH (90%) and H<sub>2</sub>O (10%). It is also purified by dissolving in a small volume of EtOH, adding a few drops of H<sub>2</sub>O and cooling. <sup>1</sup>H NMR in D<sub>2</sub>O at 44° contains 60% α-pyranose and 40% β-pyranose forms [Angyal *Adv Carbohydr Chem* **42** 15 1984.] [Smith *J Chem Soc* 1035 1940, McGeachin & Beevers *Acta Cryst* **10** 227,230 197, Beilstein **1** IV 4261.]

**D-(+)-Ribonic acid-γ-lactone** [5336-08-3] **M 148.12, m 80°, 84-86°, [α]<sub>D</sub><sup>20</sup> +18.3° (c 5, H<sub>2</sub>O).** Purify D-(+)-ribonic acid-γ-lactone by recrystallisation from EtOAc. The *tribenzoate* has **m** 54-56° (from AcOH), [α]<sub>D</sub><sup>25</sup> +27° (c 2.37, Me<sub>2</sub>NCHO), and the *3,5-O-benzylidene* derivative has **m** 230-231.5° (needles from Me<sub>2</sub>CO-petroleum ether) and [α]<sub>D</sub><sup>25</sup> -177° (CHCl<sub>3</sub>). [Chen & Joulié *J Org Chem* **49** 2168 1984, Zinner & Voigt *J Carbohydr Research* **7** 38 1968.]

**α-D(-)-Ribose** [50-69-1] **M 150.1, m 90°, [α]<sub>D</sub><sup>20</sup> -24° (after 24hours, c 10, H<sub>2</sub>O), pK<sup>25</sup> 12.22.** Crystallise α-D(-)-ribose from aqueous 80% EtOH, dry it under vacuum at 60° over P<sub>2</sub>O<sub>5</sub> and store it in a vacuum desiccator. It exhibits a complex mutarotation with : [α]<sub>D</sub><sup>10</sup> -23.1° (1.5 minutes), -21.3° (5 minutes), -19.5° (10 minutes), -19.1° (30 minutes), -21.2° (60 minutes), -23.1° (120 minutes), -23.7° (300 minutes), (c 4.5, H<sub>2</sub>O) [Phelps et al. *J Am Chem Soc* **56** 748 1934]. <sup>1</sup>H NMR in D<sub>2</sub>O at 44° shows 17% α-pyranose, 59% β-pyranose, 9% α-furanose and 15% β-furanose forms with **furanose** α-H at 5.34ppm (*J* = 3.0Hz) and β-H at 5.31 (*J* = 1.7Hz) [Angyal *Adv Carbohydr Chem* **42** 15 1984, Angyal & Pickles *Aust J Chem* **25** 1711 1972]. The *phenylhydrazone* crystallises from aqueous pyridine in yellow needles, **m** 163-164°, and the *benzylphenylhydrazone* has **m** 127-128° [Snowden *J Am Chem Soc* **72** 808 1950]. [Beilstein **1** IV 4211.]

**(+)-Rutin (quercetin-3-rubinoside)** See Vitamin P in “Miscellaneous”, this chapter.

**Saccharides.** They are separated by anion-exchange chromatography. [Walberg & Kando *Anal Biochem* **37** 320 1970.]

**D(-)-Salicin [2-(hydroxymethyl)phenyl-β-D-glucopyranoside]** [138-52-3] **M 286.3, m 204-208°, [α]<sub>D</sub><sup>25</sup> -63.5° (c ca 3, H<sub>2</sub>O).** Crystallise D(-)-salicin from EtOAc, EtOH or water and sublime it at 190-195°/12mm. [Armour et al. *J Chem Soc* 412 1961, IR: Pearl & Darling *J Org Chem* **24** 731 1959, Beilstein **17** III/IV 2986, 17/7 V 113.]

**Sennoside A (bianthraquinonyl-bis-glucoside R,R-enantiomer)** [81-27-6] **M 862.7, m 220-240°(dec), [α]<sub>D</sub><sup>20</sup> -164° (c 0.1, Me<sub>2</sub>CO/H<sub>2</sub>O 6:4), [α]<sub>D</sub><sup>20</sup> -24° (c 0.2, 70% aqueous dioxane).** Sennoside A forms yellow crystals from aqueous acetone, 2-ethoxyethanol or large volumes of H<sub>2</sub>O. [Stoll et al. *Helv Chim Acta* **32** 1896 1900, Beilstein **17** III/IV 3403.]

**Sennoside B (bianthraquinonyl-bis-glucoside R,S-enantiomer)** [128-57-4] **M 862.7, m 209-212°(dec), [α]<sub>D</sub><sup>20</sup> -100° (c 2, Me<sub>2</sub>CO/H<sub>2</sub>O 7:1), [α]<sub>D</sub><sup>20</sup> -67° (c 0.2, 70% aqueous dioxane).** Sennoside B forms yellow crystals from aqueous acetone or large volumes of H<sub>2</sub>O. [Stoll et al. *Helv Chim Acta* **32** 1896 1900, Beilstein **17** III/IV 3402.]

**Sinigrin monohydrate (Myronate K, 1-thio-β-D-glucopyranose 1-[N-(sulfooxy)-3-butenimide] monopotassium salt)** [64550-88-5] **M 415.5, m 125-127°, 127-129°, 179°(anhydrous), [α]<sub>D</sub><sup>20</sup> -17° (c 0.2, H<sub>2</sub>O), pK<sub>Est</sub> <0.** Purify sinigrin by recrystallising it three times from EtOH and once from MeOH. The *tetraacetate* has **m** 193-195°, [α]<sub>D</sub><sup>20</sup> -16° (c 0.14, H<sub>2</sub>O). [Benn et al. *J Chem Soc, Chem Commun* 445 1965, Kjaær et al. *Acta Chem Scand* **10** 432 1956, Marsh et al. *Acta Cryst (Sect B)* **26** 1030 1970.] It is a β-D-thioglucopyranoside substrate for thiogluconidase [MacLeod & Rossiter *Phytochem* **25** 1047 1986]. [Beilstein **31** H 476, **17** III/IV 3738.]

**α-Solanine (solan-5-en-3β-yl-[O<sup>3</sup>-β-D-glucopyranosyl-O<sup>2</sup>-α-L-rhamnopyranosyl-β-D-galactopyranoside])** [20562-02-1] **M 868.1, m 285°(dec), 286°(dec) (sintering >190°), [α]<sub>D</sub><sup>20</sup> -58° (c 0.8, pyridine), pK<sup>15</sup> 6.66.** Recrystallise α-solanine from EtOH, 85% aqueous EtOH, MeOH or aqueous MeOH as *dihydrate* **m** 276-278°.

Its solubility in H<sub>2</sub>O is 25mg/l and 5% in pyridine, but it is very soluble in Et<sub>2</sub>O and CHCl<sub>3</sub>. The *hydrochloride* is gummy or amorphous but has been crystallised (**m** ~212° dec). It has insecticidal properties. [Kuhn et al. *Chem Ber* **88** 1492 1955, *Beilstein* **21** III/IV 1402.]

**Solasonine (solasodine-3-O-triglycoside)** [19121-58-5] **M 884.0, m 301-303° (sinters at ~296°), [α]<sub>D</sub><sup>20</sup> -75° (c 0.5, MeOH), pK<sub>Est</sub> ~ 7.7.** Solasonine crystallises from aqueous 80% dioxane or MeOH in needles. [Bell & Briggs *J Chem Soc* **1** 1942, Briggs et al. *J Chem Soc* **4645** 1961, Briggs et al. *J Chem Soc* **2848** 1963.] The *picrate* crystallises from 30% aqueous EtOH with **m** 197-198°(dec) [Briggs & Cambie *J Chem Soc* **1422** 1958]. [*Beilstein* **27** III/IV 2006.]

**D(-)-Sorbitol (D-glucitol)** [50-70-4] **M 182.2, m 89-93° (hemihydrate), 110-111° (anhydrous), [α]<sub>D</sub><sup>20</sup> -1.8° (c 10, H<sub>2</sub>O), pK<sup>60</sup> 13.00.** Crystallise D(-)-sorbitol (as hemihydrate) several times from EtOH/water (1:1), then dry it by fusing and storing over anhydrous MgSO<sub>4</sub>. [Koch et al. *J Am Chem Soc* **75** 953 1953, *Beilstein* **1** IV 2839.]

**Starch** [9005-84-9] **M (162.1)<sub>n</sub>.** Starch is de-fatted by Soxhlet extraction with Et<sub>2</sub>O or 95% EtOH. For fractionation of starch into “amylose” and “amylopectin” fractions, see Lansky et al. [*J Am Chem Soc* **71** 4066 1949].

**Streptozotocin (N-[methylnitrosocarbamoyl]-α-D-glucosamine, streptozocin)** [18883-66-4] **M 265.2, m 111-114°(dec), 114-115°(dec), 115°(dec with evolution of gas), [α]<sub>D</sub><sup>20</sup> ~+39° (H<sub>2</sub>O, may vary due to mutarotation).** Recrystallise streptozotocin from 95% EtOH. It is soluble in H<sub>2</sub>O, MeOH and Me<sub>2</sub>CO. It has UV with λ<sub>max</sub> at 228nm (ε 6360) in EtOH. The *tetraacetate* has **m** 111-114°(dec), and [α]<sub>D</sub><sup>25</sup> +41° (c 0.78, 95% EtOH) after recrystallisation from EtOAc. [Herr et al. *J Am Chem Soc* **89** 4808 1967, NMR: Wiley et al. *J Org Chem* **44** 9 1979.] It is a potent methylating agent for DNA [Bennett & Pegg *Cancer Res* **41** 2786 1981].

**D(+)-Sucrose (β-D-fructofuranosyl-α-D-glucopyranoside)** [57-50-1] **M 342.3, m 160-186°, 186-188°, [α]<sub>D</sub><sup>20</sup> +78° (c 10, H<sub>2</sub>O), [α]<sub>D</sub><sup>20</sup> + 66° (c 26, H<sub>2</sub>O), pK<sup>25</sup> 12.62.** Crystallise D(+)-sucrose from water (solubility: 1g in 0.5ml H<sub>2</sub>O at 20°, 1g in 0.2ml in boiling H<sub>2</sub>O). It is soluble in EtOH (0.6%) and MeOH (1%). *Sucrose diacetate hexaisobutyrate* is purified by melting and, while molten, treated with NaHCO<sub>3</sub> and charcoal, then filtered. [*Beilstein* **17/8** V 399.]

**D(+)-Sucrose octaacetate** [126-14-7] **M 678.6, m 83-85°, b 260°/1mm, [α]<sub>D</sub><sup>20</sup> +71° (c 2.5, EtOH).** After three recrystallisations from EtOH or 95% EtOH (charcoal), the **m** of the *octaacetate* rises to 88-90°, or Et<sub>2</sub>O with **m** 89° and [α]<sub>D</sub><sup>25</sup> +58.5° (c 2.6, EtOH). It has a bitter taste. [Linstead et al. *J Am Chem Soc* **62** 3260 1940, Lemieux & Huber *J Am Chem Soc* **78** 4117 1956, *Beilstein* **17/8** V 410.]

**D(-)-Tagatose** [87-81-0] **M 180.2, m 131-132°, 134-135°, [α]<sub>D</sub><sup>20</sup> -6.5° (c 1, H<sub>2</sub>O).** Crystallise D(-)-tagatose from EtOH/H<sub>2</sub>O (6:1). It mutarotates from [α]<sub>D</sub><sup>22</sup> +2° (2 minutes) to -5.0° (30 minutes) (c 4, H<sub>2</sub>O). The *phenylosazone* crystallises from aqueous EtOH with **m** 185-187°(dec), and [α]<sub>D</sub><sup>23</sup> +47° (c 0.82, 2-methoxyethanol). [Totton & Lardy *J Am Chem Soc* **71** 3076 1949, Gorin et al. *Canad J Chem* **33** 1116 1955, Reichstein & Bossard *Helv Chim Acta* **17** 753 1934, Wolfrom & Bennett *J Org Chem* **30** 1284 1965, *Beilstein* **1** IV 4414.] In D<sub>2</sub>O at 27° <sup>1</sup>H NMR showed the following ratios: α-pyranose (79), β-pyranose (16), α-furanose (1) and β-furanose (4) [Angyal *Adv Carbohydr Chem* **42** 15 1984, Angyal & Pickles *Aust J Chem* **25** 1711 1972].

**Thevetin A (cardenolide glycoside)** [37933-66-7] **M 858.9, m softens at 194°, m 208-210°, [α]<sub>D</sub><sup>26</sup> -72° (c 1.48, MeOH).** Crystallise Thevetin A from H<sub>2</sub>O. The *acetyl derivative* crystallises from MeOH/Et<sub>2</sub>O at -15° with **m** 145-149°, and [α]<sub>D</sub><sup>26</sup> -54.2° (c 1.86, CHCl<sub>3</sub>). [Block et al. *Helv Chim Acta* **43** 652 1960, <sup>13</sup>C NMR: Tori et al. *Tetrahedron Lett* 717 1977, *Beilstein* **18** III/IV 2552, **18/4** V 439.]

**Thevetin B (cardenolide glycoside)** [11018-93-2] **M 858.9, m 197-201°, [α]<sub>D</sub><sup>24</sup> -61.4° (c 1.5, MeOH).** It crystallises (as *trihydrate*) from isopropanol. Dry it at 100°/0.01mm to give the *hemihydrate* (*very hygroscopic*).

[Block et al. *Helv Chim Acta* **43** 652 1960,  $^{13}\text{C}$  NMR: Tori et al. *Tetrahedron Lett* 717 1977, *Beilstein* **18** III/IV 1493.]

**$\alpha,\alpha'$ -D(+)-Trehalose (2H<sub>2</sub>O)** [6138-23-4] **M 378.3, m 96.5-97.5<sup>o</sup>, 94-100<sup>o</sup> (dihydrate), 214-216<sup>o</sup> (anhydrous),  $[\alpha]_{\text{D}}^{20} +180^{\circ}$  (dihydrate, c 4, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{20} +199^{\circ}$  (anhydrous, c 4, H<sub>2</sub>O).**  $\alpha,\alpha'$ -D(+)-Trehalose crystallises (as the *dihydrate*) from aqueous EtOH. Dry it at 13<sup>o</sup>. For the *anhydrous* compound dissolve 10g in pyridine (200ml) and distil off this solvent at atmospheric pressure, and when the temperature rises to 115.3<sup>o</sup> all the H<sub>2</sub>O is removed and 73ml of distillate is collected. Most of the anhydrous material crystallises out at this stage. The crystals are collected (6.8g), washed with Et<sub>2</sub>O to give 6.1g of anhydrous product. Higher yields are obtained by slightly more prolonged distillation. [Birch *J Chem Soc* 3489 1965, X-ray cryst: Brown et al. *Acta Cryst* **28** 3145 1972, *Beilstein* **17/8** V 3.]

**D(+)-Turanose [3-O- $\alpha$ -D-glucosido]-D-fructose** [547-25-1] **M 342.3, m 168-170<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +88^{\circ}$  (c 4, H<sub>2</sub>O).** Crystallise D(+)-turanose from H<sub>2</sub>O by addition of EtOH (its solubility is 5.3% in 95% EtOH). Form **m 157<sup>o</sup>** is obtained by crystallisation from hot MeOH, and mutarotates from +27.3<sup>o</sup> to  $[\alpha]_{\text{D}}^{20} +88^{\circ}$  (c 4, H<sub>2</sub>O). The *phenylosazone* crystallises from 15 parts of 95% EtOH with **m 200-205<sup>o</sup>**,  $[\alpha]_{\text{D}}^{20} 24.5^{\circ}$  mutarotating to +33<sup>o</sup> [24hours, c 0.82, pyridine/EtOH (4:6)]. [Pascu *Methods in Carbohydrate Chemistry I* 353 1962, Academic Press, *Beilstein* **17/7** V 213.] In D<sub>2</sub>O at 36<sup>o</sup>  $^1\text{H}$  NMR showed the following ratios:  $\alpha$ -pyranose (<4),  $\beta$ -pyranose (39),  $\alpha$ -furanose (20) and  $\beta$ -furanose (41) [Angyal *Adv Carbohydr Chem* **42** 15 1984].

**Vicine (2,4-diamino-5- $\beta$ -D-glucopyranosidoxy-6-hydroxypyrimidine)** [152-93-2] **M 304.3, m 243-244<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} -12^{\circ}$  (c 4, 0.2N NaOH).** Crystallise Vicine from water (1%) or aqueous 85% EtOH, and dry it at 135<sup>o</sup>. [Bendich & Clements *Biochim Biophys Acta* **12** 462 1953, *Beilstein* **31** H 163, **25** III/IV 4285.]

**Ustilagic acid (Ustizeain B, di-D-glucosyldihydroxyhexadecanoic acid)** [8002-36-6] **M ~780, m 146-147<sup>o</sup>,  $[\alpha]_{\text{D}}^{25} +7^{\circ}$  (c 1, pyridine),  $\text{pK}^{25} \sim 4.9$ .** Ustilagic acid is a mixture of partly acetylated di-D-glucosyldihydroxyhexadecanoic acid which crystallises from diethyl ether. It has also been purified from the culture by dissolving it in hot MeOH, filtering and concentrating by blowing a current of air until the solution becomes turbid, then heating to 50<sup>o</sup> and adding 4 volumes of H<sub>2</sub>O (also at 50<sup>o</sup>), and cooling very slowly. Filter off the white solid and dry it in air. It crystallises from Et<sub>2</sub>O, and is soluble in MeOH, butan-1,2-diol, poorly soluble in EtOH, *n*-BuOH, Me<sub>2</sub>CO and insoluble in H<sub>2</sub>O, EtOAc and \*C<sub>6</sub>H<sub>6</sub>. [Lemieux et al. *Can J Chem* **29** 409, 415 1951, Boothroyd et al. *Can J Biochem Physiol* **33** 289 1955.]

**Xanthorhamnin (xanthene rhamnoside)** [1324-63-6] **M 770.7, m 195<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +3.75^{\circ}$  (EtOH),  $\text{pK}_1^{24} 8.69$ ,  $\text{pK}_2^{24} 11.28$ ,  $\text{pK}_3^{24} 12.22$ .** Crystallise xanthorhamnin from a mixture of ethyl and isopropyl alcohols, dry it in air, then dry it further for several hours at 110<sup>o</sup>. The UV (EtOH) has  $\lambda_{\text{max}}$  at 258 and 362nm. [Nystrom et al. *J Org Chem* **22** 1272 1957, *Beilstein* **18** III/IV 3498.]

**$\alpha$ -D(+)-Xylose** [58-86-6] **M 150.1, m 146-147<sup>o</sup>, 153-154<sup>o</sup>,  $[\alpha]_{\text{D}}^{20} +92^{\circ}$  mutarotating to +18.8<sup>o</sup> (16 hours, c 10, H<sub>2</sub>O),  $\text{pK}^{18} 12.14$ .**  $\alpha$ -D(+)-Xylose forms needles or prisms (which have a very sweet taste) by slow crystallisation from aqueous 80% EtOH or absolute EtOH, which are then dried at 60<sup>o</sup> *in vacuo* over P<sub>2</sub>O<sub>5</sub>. Store it in a vacuum desiccator over CaSO<sub>4</sub>. 1Gram dissolves in 0.8ml H<sub>2</sub>O. [Bragg & Hough *J Chem Soc* 4347 1957, Hudson & Yanovsky *J Am Chem Soc* **39** 1029 1917, Monroe *J Am Chem Soc* **41** 1002 1919, *Beilstein* **1** IV 4223.] In D<sub>2</sub>O at 31<sup>o</sup>,  $^1\text{H}$  NMR showed the following ratios:  $\alpha$ -pyranose (36.5),  $\beta$ -pyranose (63),  $\alpha$ -furanose +  $\beta$ -furanose (~1) [Angyal *Adv Carbohydr Chem* **42** 15 1984, Angyal & Pickles *Aust J Chem* **25** 1711 1972].

## CAROTENOIDS

### General

Carotenoids are polyene pigments that are mostly naturally occurring in bacteria, plants and animals. They have been isolated from the natural sources and obtained first by extraction with solvents and then purified by column chromatography through  $\text{Al}_2\text{O}_3$  of various grades,  $\text{Ca}(\text{OH})_2$  alone or with  $\text{CaCO}_3$ ,  $\text{MgO}$  or Silica Gel and eluted with solvents of various polarities. The progress of separation can be followed visually because the bands of most carotenoids are of various colours. The bands can be collected by elution, or the column can be extruded and the bands cut out and extracted with a polar solvent, e.g. MeOH. This chromatography can be repeated with the separate bands, and finally the carotenoids are recrystallised to analytical purity. The purity can be checked by TLC on Silica Gel or  $\text{Al}_2\text{O}_3$  plates or paper chromatography and eluted in two dimensions. Gas-liquid or HPLC has been used for preparative work as well as for checking the purity and identifying them using internal standards such as tocopherol acetate (vitamin E acetate) and retinyl acetate.

Carotenoids are generally light sensitive, easily oxidised by air and are affected by traces of acid, e.g. in solvents. These cause the polyenes to bleach or polymerise. The necessary precautions are therefore required to minimise these effects during isolation, purification and storage. They are identified by their UV-VIS spectra, and their molar extinction coefficients at specific wavelengths ( $\lambda_{\text{max}}$ ) have been used for characterisation and for quantitation. More recently ORD, CD, NMR, IR and mass spectroscopy have been used extensively.

### Bibliography

*Carotenoids*: Karrer & Jucker, Elsevier, 1950.

*The Comparative Biochemistry of the Carotenoids*: Goodwin, Chapman & Hall Ltd, 1952.

*Carotenoids*, Vol 2, IARC Handbooks of Cancer Prevention; International Agency for Research on Cancer, 1998; ISBN 92-832-0027 and 1207-5622.

*Structures and Properties of Carotenoids in Relation to Function*: Britton, *FASEB J*, **9** 1551 1995.

*Carotenoids*: Goodwin, *Ann Rev Biochem* **24** 497 1955.

*Carotenoids*, Vol 1A, Isolation and Analysis: Britton et al. Birkhäuser Verlag, Basel, 201 1995.

*Carotenoids*, Vol 1B, Spectroscopy: Britton et al. Birkhäuser Verlag, Basel, 201 1995.

*Synthesis, Isolation and Spectra of 11 isomers of Zeaxanthin*: Englert et al. *Helv Chim Acta* **74** 969 1992.

*Chromtography*: Kushwaha et al. *J Biol Chem* **245** 4708 1970.

*HPLC*: Khachik et al. *Anal Chem* **69** 1873 1997; Khachik et al. *Pure & Appl Chem* **63** 71 1996.

*UV-VIS spectra*: Zechmaster *Experientia* **10** 1 1954.

*Classification of CD Spectra of Carotenoids*: Sturzenegger et al. *Helv Chim Acta* **63** 1074 1980.

*ORD of Carotenoids*: Bartlett et al. *J Chem Soc (C)* 2527 1969.

*Carotenoids, Physical Chemical and Biological Functions and Properties*: J.T.Lendrum, Taylor & Francis Ltd, 2008; ISBN 9781420052305.

*Carotenoids Handbook*: H.P. Pfander, G. Britton and L-J. Synnove, Birkhauser Verlag AG, 2004; ISBN 9783764361808.

*Carotenoids*, Vol 5, Nutrition and Health, G. Britton, L-J. Synnove and H. Pfander, Birkhauser Verlag AG, 2008; ISBN 0783764375003.

**Antheraxanthin (5,6-epoxy-zeaxanthin)** [68831-78-7] **M 584.8, m 205°,  $\lambda_{\text{max}}$  460.5, 490.5nm, in  $\text{CHCl}_3$ .**

Likely impurities are violaxanthin (see later) and mutatoxanthin (which contains a dihydrofuran ring derived from the epoxy group of antheraxanthin). Purify it by chromatography on columns of  $\text{Ca}(\text{OH})_2$  and of  $\text{ZnCO}_3$ . It crystallises from  $^*\text{C}_6\text{H}_6/\text{MeOH}$  as needles or thin plates. Store it in the dark, under  $\text{N}_2$  or argon at  $-20^\circ$ . In  $\text{CHCl}_3/\text{SbCl}_2$  it gives a blue colour with  $\lambda_{\text{max}}$  at 587nm and gives a stable blue colour with conc HCl. [Karrer & Oswald *Helv Chim Acta* **18** 1303 1935, Karrer & Jucker *Helv Chim Acta* **28** 300 1945.]

**$\beta$ -Apo-4'-carotenal** [12676-20-9] **M 414.7, m 139°,  $A_{1\text{cm}}^{1\%}$  2640 at 461nm.** Recrystallise the carotenal from  $\text{CHCl}_3/\text{EtOH}$  mixture or *n*-hexane. The  $\lambda_{\text{max}}$  of the absorption spectra of the radical anions and cations of such polyenals formed by pulse radiolysis are linearly dependent on the number of conjugated double bonds [Land et al. *JCS, Faraday Trans 1* **74** 538 1978, Bobrowski & Das *J Phys Chem* **91** 1210 1987, *Beilstein* **7** IV 1782.]

**trans- $\beta$ -Apo-8'-carotenal** [1107-26-2] M 414.7, m 136-139°. Recrystallise  $\beta$ -apo-8'-carotenal from  $\text{CHCl}_3/\text{EtOH}$  mixture or *n*-hexane. The absorption spectra of the radical anion of this  $\text{C}_{30}$  aldehyde in CTAB and Triton X100 micelles suggested an alcohol-like nature of the environment probed by the long-chain polyenal. For effect of conjugation on spectra see preceding entry [Bobrowski & Das *J Phys Chem* 91 1210 1987, Land et al. *JCS, Faraday Trans 1* 74 538 1978, *Beilstein* 7 III 2622, 7 IV 1782.]

**$\beta$ -Apo-8'-carotenoic acid ethyl ester** [1109-11-1] M 526.8, m 134-138°,  $A_{1\text{cm}}^{1\%}$  2550 at 449nm. Recrystallise the ester from petroleum ether or petroleum ether/EtOAc. Store it in the dark under  $\text{N}_2$  or argon at  $-20^\circ$ . [*Beilstein* 9 IV 2701 for Me ester.]

**$\beta$ -Apo-8'-carotenoic acid methyl ester** [16266-99-2] M 512.7, m 136-137°,  $A_{1\text{cm}}^{1\%}$  2575 at 446nm and 2160 at 471nm, in petroleum ether. Recrystallise the methyl ester from petroleum ether or petroleum ether/EtOAc. Store it in the dark in an inert atmosphere at  $-20^\circ$ . [*Beilstein* 9 IV 2701.]

**Astacin ( $\beta,\beta$ -carotene-3,3',4,4'-tetraone, Astacene)** [514-76-1] M 592.8, m 228°, 232-233°, 240-243°(evacuated tube),  $\epsilon_{1\text{cm}}^{1\%}$  550,000 at 498nm (pyridine). Astacin is a red pigment found in marine organisms (protozoa, algae, crustacians, sponges, fish, reptiles) together with astaxanthin (see below) from which it is possibly formed by autoxidation [Kuhn & Lederer *Chem Ber* 66 488 1933]. A probable impurity is astaxanthin (3,3'-dihydroxy- $\beta,\beta$ -carotene-4,4'-dione) which may be its precursor. It has been prepared by adding a solution of canthaxanthene (1.0g, see [514-78-3] below) in  $^*\text{C}_6\text{H}_6$  (15ml) to *t*-BuOK in *t*-BuOH (1.40N, 130ml) and shaken with  $\text{O}_2$  at  $20^\circ$  for 30 hours. Water (200ml) is added, then 0.5N HCl (800ml) and extracted with  $\text{CHCl}_3$  (300ml). The extract is washed with aqueous  $\text{NaHCO}_3$  (3 x 300ml),  $\text{H}_2\text{O}$  (2 x 200ml), diluted with  $^*\text{C}_6\text{H}_6$  (50ml) and the organic layer is evaporated *in vacuo*. The residue is recrystallised from  $\text{CHCl}_3$ -EtOH (1:5) to give *astacin* (450mg, ~90%) as deep purple leaflets m 232-233° with  $\lambda_{\text{max}}$  (pyridine) 498mm ( $\epsilon$  100 x 10<sup>3</sup>), ( $\text{CS}_2$ ) 513mm, (hexane) 477mm, (EtOH) 483mm, and (KOH-EtOH) 478mm; the IR ( $\text{CHCl}_3$ ) has  $\nu_{\text{max}}$  at 3410br ( $\epsilon$  ~90), 1610 ( $\epsilon$  1200), 1550, 1327 and 969 ( $\epsilon$  895)  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.30 (12H, C-1  $\text{Me}_2$  and C-1'  $\text{Me}_2$ ), 2.10 (6H, C-5 Me and C-5' Me), 2.02 (6H, C-9 Me and C-9' Me), 2.02 (6H, C-13 Me and C-13' Me), and 6.06 (C-2 H and C-2' H) (14 olefinic H-7,7', 8,8', 10,10', 11,11', 12,12', 14,14', and 15,15' are not included); and  $m/z$  592  $\text{M}^+$ . [Cooper et al. *J C S Perkin Trans I* 2195 1975.] Alternatively, purify astacin by chromatography on alumina/fibrous clay (1:4) or mixed chromatograms on sucrose using  $^*\text{C}_6\text{H}_6$ /light petroleum 1:4 as eluant; or by partition between petroleum ether and MeOH (alkaline). It crystallises from pyridine/water in purple leaflets or needles with a metallic lustre. Store it in the dark under  $\text{N}_2$  at  $-20^\circ$ . It is very soluble in pyridine, dioxane,  $\text{CS}_2$ ,  $\text{CHCl}_3$  and dilute aqueous NaOH, slightly soluble in AcOH, EtOAc and  $^*\text{C}_6\text{H}_6$ , but almost insoluble in  $\text{H}_2\text{O}$ ,  $\text{Et}_2\text{O}$ , petroleum ether and MeOH. Its *diacetate (enol-acetate)*, obtained by  $\text{Ac}_2\text{O}$ /pyridine treatment at  $25^\circ$  for 34 hours has m 232-233° (dec) (from  $\text{CHCl}_3/\text{EtOH}$ , 1:5) or 235° (dec) (from pyridine +  $\text{H}_2\text{O}$ ), and forms violet/black needles with  $\lambda_{\text{max}}$  (pyridine) at 497mm ( $\epsilon$  109 x 10<sup>3</sup>) and (EtOH) 474mm; the IR ( $\text{CHCl}_3$ ) has  $\nu_{\text{max}}$  at 1756 ( $\epsilon$  530), 1639 ( $\epsilon$  1150), 1555 and 972  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 1.35 (12H, C-1  $\text{Me}_2$  and C-1'  $\text{Me}_2$ ), 2.04 (6H, C-5 Me and C-5' Me), 2.04 (6H, C-9 Me and C-9' Me), 2.02 (6H, C-13 Me and C-13' Me), and 2.29 (2 OAc) (14 olefinic H-7,7', 8,8', 10,10', 11,11', 12,12', 14,14', and 15,15' are not included); and  $m/z$  576  $\text{M}^+$ . The *dipalmitate (enol-ester, Astacein)* has m 121°, and forms red square leaflets from petroleum ether. [Davis & Weedon *Proc Chem Soc* 182 1960, Widmer et al. *Helv Chim Acta* 65 671 1982, Karrer et al. *Helv Chim Acta* 17 412,745 1934, 18 96 1935, 19 479 1936, *Beilstein* H 30 102, 7 III 4797.]

**Astaxanthin (3,3'-dihydroxy- $\beta,\beta$ -carotene-4,4'-dione)** [472-61-7] M 596.84, m 182-183°. It is a potent antioxidant from marine alge, red yeast and other plant and marine animal sources. It was isolated from lobster eggs, is present in salmon, the red feathers of some birds (*Laniarius* spp, and flamingo), and occurs in flower petals of some *Ranunculaceae*. Racemic astaxanthin has been synthesised from astacene (see above) by reduction with  $\text{KBH}_4$  to  $\beta,\beta$ -carotene-3,4,3',4'-tetraol followed by oxidation with 2,3-dichloro-5,6-dicyano-*p*-benzoquinone and purification by preparative TLC on Kieselgel (20% $\text{Me}_2\text{CO}$ /light petroleum) to give ( $\pm$ )-*astaxanthin* m 182-183° as needles, after crystallisation from  $\text{Me}_2\text{CO}$ /light petroleum. It has  $\lambda_{\text{max}}$  (MeOH)

472mm, (hexane) 466-467mm, (CHCl<sub>3</sub>) 485 and (CS<sub>2</sub>) 503mm; the IR (CHCl<sub>3</sub>) has  $\nu_{\max}$  at 3620, 3510, 1660, 1610 and 975 cm<sup>-1</sup>; the <sup>1</sup>H NMR (60MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 1.31 (6H, C-1 Me and C-1' Me), 1.20 (6H, C-1 Me and C-1' Me), 1.94 (6H, C-5 Me and C-5' Me), 1.98 (12H, C-9 Me, C-9' Me, C-13 Me and C-13' Me), (14 olefinic H-7,7', 8,8', 10,10', 11,11', 12,12', 14,14', and 15,15' are not included); and  $m/z$  592.386 M<sup>+</sup>. TLC in the above system or Micro-Cel did not separate it from natural (3*S*,3'*S*)-astaxanthin from the common lobster or sea water crayfish. Synthetic ( $\pm$ )-astaxanthin combined with the appropriate apoprotein to form  $\alpha$ -crustacyanin with  $\lambda_{\max}$  (MeOH) at 630mm. Its diacetate has **m 203-205**<sup>o</sup> (blue-black needles from pyridine/H<sub>2</sub>O, **m 189-191**<sup>o</sup> also reported),  $\lambda_{\max}$  (CHCl<sub>3</sub>) with 482mm, and is also indistinguishable by mixed TLC on Kieselgel, Micro-Cel or alumina from natural (3*S*,3'*S*)-astaxanthindiacetate, **m 187-189**<sup>o</sup>. All manipulations have to be carried out under argon to avoid their ready oxidation to astacin (see above). [Cooper et al *J C S Perkin Trans I* 2195 1975, Kienzel & Mayer *Acta Chim Acta* **61** 2609 1978.]

An efficient synthesis of natural (3*S*,3'*S*)-astaxanthin from (4*R*,4'*R*)-4-hydroxy-2,2,6-trimethylcyclohexanone in 67% yield and high optical purity has been described and was purified by recrystallisation from CH<sub>2</sub>Cl<sub>2</sub>/MeOH to give dark violet crystals, **m 216-218**<sup>o</sup>, with R<sub>F</sub> 0.30 on TLC with kieselgel (CH<sub>2</sub>Cl<sub>2</sub>/Me<sub>2</sub>CO/HCOOH, 95:5:3); its UV has  $\lambda_{\max}$  nm ( $\epsilon$ ), (CHCl<sub>3</sub>) at 492 (1920), 488 (125,000), and (CH<sub>2</sub>Cl<sub>2</sub>) 324(14,500), 298 (15,500), ~285sh (~14,000), 251 (14,500), ~225 (~15,000); and IR (KBr) has  $\nu_{\max}$  at 3486m (OH), 1664s (C=O, conjugated), 1607m, 1557s (C=O, conjugated), 1399w, 1390m, 1366w (gem Me<sub>2</sub>), 1076s, 1039m (OH) and 969s (CH=CH, *trans*) cm<sup>-1</sup>. [Widmer et al. *Helv Chim Acta* **64** 2405 1981]. Other high yielding syntheses have been achieved [Becher et al. *Helv Chim Acta* **64** 2419 1981, Kienzel & Mayer *Helv Chim Acta* **61** 2609 1978], and the absolute configuration of natural (3*S*,3'*S*)-astaxanthin, **m 219-220**<sup>o</sup> has been determined by X-ray analysis of a crystalline synthetic intermediate [Widmer et al. *Helv Chim Acta* **64** 2405 1981] and comparison of CD spectra [Englert et al. *Helv Chim Acta* **60** 1209 1977]. Its <sup>1</sup>H NMR (270MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  for olefinic protons at 6.21 (H-7 and H-7'), 6.43 (H-8 and H-8'), 6.30 (H-10 and H-10'), 6.66 (H-11 and H-11'), 6.45 (H-12 and H-12'), ~6.30 (H-14 and H-14') and ~6.68 (H-15 and H-15') with  $J_{7,8}$  or  $J_{7,8'}$  = 16.3Hz,  $J_{10,11}$  or  $J_{10,11'}$  = 11.5Hz and  $J_{11,12}$  or  $J_{11',12'}$  = 14.7Hz;  $\delta$  for Me groups at 1.21 and 1.32 (6H, C-1 Me and C-1' Me), 1.95 (6H, C-5 Me and C-5' Me), 2.00 (6H, C-9 Me and C-9' Me), 1.99 (6H, C-13 Me and C-13' Me); and  $\delta$  for cyclohexenone protons at 1.82 (2H, axial H-2 and H-2'), 2.16 (2H, equatorial H-2 and H-2'), 4.32 (2H, H-3 and H-3'), and 3.70 (2H, 3-OH and 3'-OH,  $J$  = 1.8Hz,) with  $J_{2,2}$  (gem) or  $J_{2',2'}$  (gem) = 12.7Hz,  $J_{2e,3a}$  (vic) = 5.7Hz and  $J_{2a',3a'}$  (vic) = 14.0Hz,  $J_{2e',3a'}$  (vic) = 5.7Hz and  $J_{2a',3a'}$  (vic) = 14.0Hz; the <sup>13</sup>C NMR (23MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  for methyl group carbons at 12.8 (13,13'), 12.6 (9,9'), 13.9 (5,5'), 26.2 (1 or 1'), 30.8 (1' or 1); for end group carbons 36.8 (1,1'), 45.6 (2,2'), 69.2 (3,3'), 200.3 (4,4'), 126.9 (5,5'), 162.4 (6,6'); and for olefinic carbons at 123.2 (7,7'), 142.2 (8,8'), 134.7 (9,9'), 135.1 (10,10'), 124.6 (11,11'), 139.7 (12,12'), 136.7 (13,13'), 133.8 (14,14') and 130.7 (15,15'); and the CD spectrum in CH<sub>2</sub>Cl<sub>2</sub> has  $\lambda_{\max}$  ( $\Delta\epsilon$ ) at 224 (+12.8), 249 (-14.4), 280 (+ 12.5), 323 (-23.1), 384 (+6.7) and 521 (-3.2, very broad) [Englert et al. *Helv Chim Acta* **60** 1209 1977].

It is the prosthetic group in many carotenoproteins, mainly in invertebrates, which vary from colourless to bright blue. It reduces blood glucose and improves some parameters in diabetic metabolism. It ameliorates blood flow and vascular tone in hypertension model and regulates *connexin* 43 in *in vitro* studies, and may be a cancer chemopreventive drug. It is obtained on a large scale from the heterobasidiomycetous yeast *Phaffia rhodozyma* and plays an important role in carotene biosynthesis [Schroeder & Johnson *J Biol Chem* **270** 18374 1995].

**Bixin (6,6'-diapo- $\psi,\psi$ -carotenedioic acid monomethyl ester)** [6983-79-5] **M 394.5, m 198<sup>o</sup>, 217<sup>o</sup>(dec)**,  $\lambda_{\max}$  (CHCl<sub>3</sub>) **209, 475 and 443nm**, **pK<sub>Est</sub> ~4.3**. Crystallise bixin from Me<sub>2</sub>CO (violet prisms) [Pattenden et al. *J Chem Soc (C)* 235 1970]. [Beilstein **2** III 2020, **2** IV 2455, H **30** 110.]

**Canthaxanthin (trans) ( $\beta$ ,  $\beta$ -carotene-4,4'-dione)** [514-78-3] **M 564.9, m 211-212<sup>o</sup>, 213<sup>o</sup>,  $\epsilon_{\text{mol}}$  124,100 (at 466nm, petroleum ether), 118,000 (at 484nm \*C<sub>6</sub>H<sub>6</sub>)**. Purify canthaxanthin by chromatography on a column of deactivated alumina or magnesium oxide, or on a thin layer plate of silica gel G (Merck), using dichloromethane/diethyl ether (9:1) to develop the chromatogram. Crystallise it from CH<sub>2</sub>Cl<sub>2</sub> or CH<sub>2</sub>Cl<sub>2</sub>/MeOH. Store in the dark and in an inert atmosphere at -20<sup>o</sup>. [Beilstein **7** III 4364, **7** IV 2680.]

**Capsanthin [(3*R*,3'*S*,5'*R*)-3,3'-dihydroxy- $\beta,\kappa$ -carotene-6-one]** [465-42-9] **M 584.9, m 170<sup>o</sup>,  $\lambda_{\max}$  484 ( $\epsilon_{\text{mol}}$  121,000), inflexion at 515 nm, in hexane**. Possible impurities are zeaxanthin and capsorubin (below). Purify

capsanthin by chromatography on a column of alumina (grade IV) and develop with 0.2% EtOH in \*C<sub>6</sub>H<sub>6</sub>. It crystallises from petroleum ether in red needles, IR has  $\nu_{\max}$  at 3600, 1664cm<sup>-1</sup>, and  $\epsilon_{1033}/\epsilon_{1664} \sim 4.1$ . [Baker et al. *J Chem Soc* 4019 1961, ORD & absolute config: Bartlett et al. *J Chem Soc (C)* 2527 1969.] The *di-O-acetate* is purified on an alumina (grade II) column in \*C<sub>6</sub>H<sub>6</sub> and it is recrystallised from MeOH with **m** 160-162°. [Faigle & Karrer *Helv Chim Acta* 44 1257 1961, *Beilstein* 8 II 415, 8 III 3047, 8 IV 2657.]

**Capsorubin (3,3'-dihydroxy- $\kappa,\kappa$ -carotene-6,6'dione)** [470-38-2] **M 604.9, m 218°,  $\lambda_{\max}$  443, 468, 503 nm, in hexane.** Possible impurities are zeaxanthin and capsanthin. Purify capsorubin by chromatography on a column of CaCO<sub>3</sub> or MgO. Crystallise it from \*benzene/petroleum ether or CS<sub>2</sub>. [*Beilstein* 1 III 3327, 8 III 3873, 8 IV 3304.]

**$\alpha$ -Carotene (6'R- $\alpha$ -carotene)** [7488-99-5] **M 536.9, m 184-188°, 187.5° (evacuated tube),  $[\alpha]_D^{20} +385^\circ$  (c 0.08, \*C<sub>6</sub>H<sub>6</sub>),  $[\alpha]_D^{25} +538^\circ$ ,  $\lambda_{\max}$  422, 446, 474 nm, in hexane,  $A_{1\text{cm}}^{1\%}$  2725 (at 446nm), 2490 (at 474nm),  $\epsilon_{\text{mol}}$  145,300 (at 455nm, hexane), and 456 (at 485nm, \*C<sub>6</sub>H<sub>6</sub>).** Purify  $\alpha$ -carotene by chromatography on columns of calcium hydroxide, alumina or magnesia. Crystallise it from CS<sub>2</sub>/MeOH, toluene/MeOH, diethyl ether/petroleum ether, or acetone/petroleum ether. Store it in the dark, under N<sub>2</sub> or Ar at -20°. It gives a blue colour with  $\lambda_{\max}$  at 542nm when mixed with SbCl<sub>3</sub> in CHCl<sub>3</sub>. [Karrer & Walker *Helv Chim Acta* 16 641 1933, Eugster et al. *Helv Chim Acta* 52 1729 1969, Eugster & Karrer *Helv Chim Acta* 38 610 1955, Strain *J Biol Chem* 105 523 1934, *Beilstein* 5 III 2457, 5 IV 2620.]

**all-trans- $\beta$ -Carotene** [7235-40-7] **M 536.9, m 178-179°, 179-180°, 180°, 181°, 183° (evacuated capillary),  $\epsilon_{\text{mol}}$  138,900 (at 450nm, petroleum ether), 124,300 (at 462nm, \*C<sub>6</sub>H<sub>6</sub>).** It forms purple prisms when crystallised from \*C<sub>6</sub>H<sub>6</sub>/MeOH and red rhombs from petroleum ether. Its solubility in hexane is 0.1% at 0°. It is **oxygen sensitive** and should be stored under N<sub>2</sub> at -20° in the dark. It gives a deep blue colour with  $\lambda_{\max}$  at 590nm when mixed with SbCl<sub>3</sub> in CHCl<sub>3</sub>. Its UV (\*C<sub>6</sub>H<sub>6</sub>) has  $\lambda_{\max}$  at 429infl, 454 and 484nm. The principal peak at 454nm has  $A_{1\text{cm}}^{1\%}$  2000. [Synthesis: Surmatis & Ofner *J Org Chem* 26 1171 1961; Milas et al. *J Am Chem Soc* 72 4844 1950.]  $\beta$ -Carotene is also purified by column chromatography (Al<sub>2</sub>O<sub>3</sub> activity I-II). It is dissolved in petroleum ether/\*C<sub>6</sub>H<sub>6</sub> (10:1), applied to the column and eluted with petroleum ether/EtOH; the desired fraction is evaporated and the residue is recrystallised from \*C<sub>6</sub>H<sub>6</sub>/MeOH (violet-red plates). [UV: Inhoffen et al. *Justus Liebigs Ann Chem* 570 54, 68 1950; Review: Fleming *Selected Organic Synthesis* (J Wiley, Lond) pp. 70-74 1973.] *Alternatively*, it can be purified by chromatography on a magnesia column, thin layer of Kieselguhr or magnesia. Crystallise it from CS<sub>2</sub>/MeOH, Et<sub>2</sub>O/petroleum ether, acetone/petroleum ether or toluene/MeOH. Store it in the dark, under an inert atmosphere, at -20°. Recrystallise it also from 1:1 EtOH/CHCl<sub>3</sub>. [Bobrowski & Das *J Phys Chem* 89 5079 1985, Johnston & Scaiano *J Am Chem Soc* 108 2349 1986, Strain *J Biol Chem* 105 523 1934, *Meth Biochem Anal* 4 1 1957, *Beilstein* 5 II 638, 5 III 2453, 5 IV 2617.]

**$\delta$ -Carotene (6R- $\epsilon,\psi$ -carotene)** [472-92-4] **M 536.9, m 140.4°, 142° (evacuated capillary)  $[\alpha]_D^{25} +352^\circ$  (c 16 hexane),  $[\alpha]_{\text{Cd}} +317^\circ$  (CS<sub>2</sub>),  $\lambda_{\max}$  430, 456, 486nm (hexane).**  $\delta$ -Carotene crystallises from CS<sub>2</sub>/hexane/EtOH as red needles. The **racemic** carotene is purified through an alumina (Grade II) column by elution with 15% \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 60-80°), and the main band eluent is evaporated and the residue is crystallised from MeOH/petroleum ether (b 60-80°) to give  $\delta$ -carotene with **m** 150-151°. [Porter & Anderson *Arch Biochem* 32 21 1951, Synthesis: Manchand et al. *J Chem Soc* 2019 1965, Absolute Config: Buchecker & Eugster *Helv Chim Acta* 54 327 1971, 5 III 2453, 5 IV 2617.]

**$\gamma$ -Carotene ( $\beta,\psi$ -carotene)** [472-93-5, 10593-83-6] **M 536.9, m 152-153.5° (synthetic), 177.5° (polymorph),  $A_{1\text{m}}^{1\%}(\lambda_{\max})$  2055 (at 437nm), 3100 (at 462nm), 2720 (at 494nm) in hexane.** Purify  $\gamma$ -carotene by chromatography on alumina [Grade II in petroleum ether (b 60-80°) and elute with \*C<sub>6</sub>H<sub>6</sub>], or magnesia columns. When crystallised from \*C<sub>6</sub>H<sub>6</sub>/MeOH (2:1), it had **m** 177.5°. Store it in the dark, under an inert atmosphere at 0°, or in an evacuated tube at -20°. The purity is verified by TLC on Ca(OH)<sub>2</sub>/Kieselgel (8:2) using petroleum ether (b 60-80°) as eluant. [Manchand et al. *J Chem Soc* 2019 1965, *Beilstein* 5 III 2453, 5 IV 2617.]

**$\xi$ -Carotene** [38894-81-4] **M 536.9, m 38-42°,  $\lambda_{\max}$  378, 400, 425nm,  $A_{1\text{m}}^{1\%}(\lambda_{\max})$  2270 (400nm), in petroleum ether.** Purify  $\xi$ -carotene by chromatography on 50% magnesia-HyfloSupercel, developing with hexane and

eluting with 10% EtOH in hexane. It crystallises from toluene/MeOH. [Gorman et al. *J Am Chem Soc* **107** 4404 1985.] Store it in the dark under N<sub>2</sub> or argon at -20°. Also purify it like  $\gamma$ -Carotene. [Beilstein **5** IV 2623.]

**Citraxanthin** [5,9,14,18-tetramethyl-20-(2,6,6-trimethyl-1-cyclohex-1-enyl)- 3,5,7,9, 11,13,15,17,19-eicosanonen-2-one] [3604-90-8] **M 456.7, m 155-156°,  $A_{1m}^{1\%}(\lambda_{max})$  410 (349nm), 275 (466nm) in hexane.** Purify it by chromatography on a column of 1:1 MgO and HyfloSupercel (diatomaceous filter aid). Crystallise it from petroleum ether. Store it in the dark under N<sub>2</sub> or argon at 0°.

**Crocetin diethyl ester (8,8'-diapo- $\psi,\psi$ -carotenedioic acid diethyl ester)** [5056-14-4] **M 384.5, m 218-219°, 222.5°,  $A_{1m}^{1\%}(\lambda_{max})$  2340 (at 400nm), 3820 (at 422nm), 3850 (at 450nm) in petroleum ether.** Purify the diethyl ester by chromatography on a column of silica gel G. Recrystallise it from \*benzene. Store it in the dark, under N<sub>2</sub> or argon, at 0°. [Beilstein **2** III 2018, H **30** 106.]

**$\alpha$ -Cryptoxanthin (all-trans-3R,6R'- $\beta$ ,  $\epsilon$ -carotene-3-ol, zeinoxanthin, physoxanthin)** [24480-38-4] **M 552.9, m 175-176° (racemate 157.5-158.5°),  $[\alpha]_D^{25}$  -508° (c 0.4 Me<sub>2</sub>CO), CD<sub>max</sub>  $\Delta\epsilon$  -3.6 (218nm), +4.4 (245nm) and -9.0 (283nm),  $\lambda_{max}$  (hexane) 421, 446, 474nm,  $\epsilon_{mol}$  145,500 (at 446nm, hexane), 130,000 (at 459nm, \*C<sub>6</sub>H<sub>6</sub>).** It crystallises from \*C<sub>6</sub>H<sub>6</sub>/MeOH in red needles. The *racemate* is purified through an Al<sub>2</sub>O<sub>3</sub> column, eluting with \*C<sub>6</sub>H<sub>6</sub> and recrystallising from MeOH/Et<sub>2</sub>O. It has  $\lambda_{max}$  ( $\epsilon$  10<sup>-3</sup>) (hexane) at 421, 444, 471nm (90, 133, 122). [Loeber et al. *J Chem Soc (C)* 404 1971, Goodfellow et al. *J Chem Soc, Chem Commun* 1578 1970, Beilstein **6** IV 5113.]

**$\beta$ -Cryptoxanthin (all-trans-3R- $\beta,\beta$ -carotene-3-ol, caricaxanthin)** [472-70-8] **M 552.9, m 169° (natural), (racemate m 172-173°),  $\epsilon_{mol}$  131,900 (at 449nm, petroleum ether).** Purify it by chromatography on MgO, CaCO<sub>3</sub> or deactivated alumina, using EtOH or diethyl ether to develop the column. Crystallise it from \*C<sub>6</sub>H<sub>6</sub>/EtOH (metallic prisms), or needles from \*C<sub>6</sub>H<sub>6</sub>. Store it in the dark under N<sub>2</sub> or argon at -20°. The *acetate* has **m 117.5°**. The *racemate* is purified through a column of alumina (grade IV), eluted with \*C<sub>6</sub>H<sub>6</sub> then EtOAc/\*C<sub>6</sub>H<sub>6</sub> (1:9) and recrystallised from petroleum ether (b 60-80°) with **m 172-173°**. [Loeber et al. *J Chem Soc (C)* 404 1971, Goodfellow et al. *J Chem Soc Chem Commun* 1578 1970, Isler et al. *Helv Chim Acta* **40** 456 1957, Beilstein **6** III 3772, **6** IV 5111.]

**3,30-Diketospirilloxanthin (all-trans-2,31-dimethoxy-2,6,10,14,19,23,27,31-octamethyl-4,6,8,10,12,-14,16,18,20,22,24,26,28-dotriacontatridecene-3,30-dione)** [24009-17-4] **M 624.9, m 225-227°,  $\epsilon_{1m}^{1\%}(\lambda_{max})$  550 (at 349nm), 820 (at 422nm), 2125 (at 488nm), 2725(516nm), 2130(at 551nm) in hexane.** Purify it by chromatography on a column of partially deactivated alumina. Recrystallise it from acetone/petroleum ether. Store it in the dark, in an inert atmosphere at 0°. [cf Beilstein **1** III 2297, **1** IV 2750.]

**Echinenone ( $\beta,\beta$ -caroten-4-one)** [432-68-8] **M 550.8, m 178-179°,  $A_{1m}^{1\%}(\lambda_{max})$  2160 (at 458nm) in petroleum ether.** Purify  $\beta,\beta$ -caroten-4-one by chromatography on partially deactivated alumina or magnesia, or by using a thin layer plate of silica gel G with 4:1 cyclohexane/diethyl ether as the developing solvent. Recrystallise it from \*C<sub>6</sub>H<sub>6</sub>/MeOH. Store it in the dark at -20°. The *oxime* crystallises from \*C<sub>6</sub>H<sub>6</sub> with **m 208°**. [Beilstein **7** III 2858, **7** IV 1881.]

**Ethyl bixin (6,6'-diapo- $\Psi,\Psi$ -carotenedioic acid monomethyl ester monoethyl ester)** [6895-43-8] **M 436.6, m 138°.** Crystallise the ester from EtOH. [cf. Beilstein **2** III 2020, **2** IV 2356.]

**Lutein** See xanthophyll.

**Lycopene (all-trans- $\Psi,\Psi$ -carotene)** [502-65-8] **M 536.9, m 172-173°, 174°,  $\epsilon_{mol}$  184,900 (470nm petroleum ether), 180,600 (at 487nm \*C<sub>6</sub>H<sub>6</sub>).** Crystallise lycopene from CS<sub>2</sub>/MeOH, diethyl ether/petroleum ether, or

acetone/petroleum ether. Also purify it by column chromatography on deactivated alumina, CaCO<sub>3</sub>, calcium hydroxide or magnesia. It is oxygen sensitive and is stored in the dark, in an inert atmosphere. Also purified like  $\gamma$ -Carotene. [*Beilstein* 1 III 1076, 1 IV 1165.]

**Lycoxanthin (all-trans- $\Psi$ , $\Psi$ -carotene-16,16'-ol)** [19891-74-8] M 268.3, m 173-174°,  $\epsilon_{1\text{cm}}^{1\%}$  3360 (472.5nm), also  $\lambda_{\text{max}}$  444 and 503nm in petroleum ether. Crystallise lycoxanthin from diethyl ether/light petroleum, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or CS<sub>2</sub>. Purify it also by chromatography on columns of CaCO<sub>3</sub>, Ca(OH)<sub>2</sub> or deactivated alumina, and washing with \*benzene and eluting with 3:1 \*C<sub>6</sub>H<sub>6</sub>/MeOH. Store it in the dark, in an inert atmosphere, at -20°. The *dipalmitate ester* crystallises from \*C<sub>6</sub>H<sub>6</sub>/MeOH and has m 76°. [*Beilstein* 1 III 2051, 1 IV 2368.]

**Methylbixin (6,6'-diapo- $\Psi$ , $\Psi$ -carotenedioic acid dimethyl ester)** [26585-94-4] M 408.5, m stable *trans* form 203°, 205-206° (corr), unstable *cis* form 164°,  $\lambda_{\text{max}}$  405, 425, 450, 484nm in hexane. Crystallise the dimethyl ester from EtOH/CHCl<sub>3</sub>, or \*benzene. Also purify it by chromatography on alumina (Grade III) and eluting with 9:1 \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 60-80°), and recrystallise it from EtOAc. [Pattenden et al. *J Chem Soc (C)* 235 1970, *Beilstein* 2 III 2020, 2 IV 2356.]

**Physalien (all-trans  $\beta$ -carotene-3,3'(R,R)-diol dipalmitate)** [144-67-2] M 1044, m 98.5-99.5°,  $A_{1\text{cm}}^{1\%}(\lambda_{\text{max}})$  1410 (at 449nm), 1255 (at 478nm) in hexane. Purify it by chromatography on water-deactivated alumina, using hexane/diethyl ether (19:1) to develop the column. It crystallises from \*benzene/EtOH in red needles. Store in the dark, in an inert atmosphere, at 0°. [See Zeaxanthin dipalmitate.] [*Beilstein* 6 III 5970.]

**Phytoene (7,7',8,8',11,11',12,12'-octahydro- $\psi$ , $\psi$ -carotene)** [540-04-5] M 544.9,  $A_{1\text{cm}}^{1\%}(\lambda_{\text{max}})$  850 (at 287nm) in hexane,  $\lambda_{\text{max}}$  275, 287 and 297nm. Purify phytoene by chromatography on columns of magnesium oxide-Supercel (a diatomaceous filter aid) or alumina [Rabourn et al. *Arch Biochem Biophys* 48 267 1954]. Store it as a solution in petroleum ether under nitrogen at -20°. [*Beilstein* 1 IV 1155.]

**Phytofluene (all-trans- 5,6,7,8,9,10,10',9',8',7',6',5'-dodecahydrolycopene)** [540-05-6] M 549.0, b 140-185°(bath temperature)/0.0001mm,  $A_{1\text{cm}}^{1\%}(\lambda_{\text{max}})$  1350 (at 348nm) in petroleum ether,  $\lambda_{\text{max}}$  331, 348, 267. Purify it by chromatography on partially deactivated alumina [Goodwin *Biochem J* 53 538 1953, Kushwaha et al. *J Biol Chem* 245 4708 1970]. Store it as a solution in petroleum ether under N<sub>2</sub> at -20°. [*Beilstein* 1 III 1072, 1 IV 1159.]

**Polycopene (all-cis [Z]- $\psi$ , $\psi$ -carotene or lycopene)** [2361-24-2] M 536.5, m 111°,  $\lambda_{\text{max}}$  443.5, 470nm in petroleum ether. Purify polycopene by chromatography on deactivated alumina [Kushwaha et al. *J Biol Chem* 245 4708 1970]. Crystallise it from petroleum ether. Store it in the dark, under N<sub>2</sub> or argon at -20°. [*Beilstein* 1 III 1079, 1 IV 1167.]

**Proneurosporene (3,4,7',8'-tetrahydrolycopene)** [10467-46-6] M 538.9,  $\lambda_{\text{max}}$  408, 432, 461 nm,  $\epsilon_{1\text{cm}}^{1\%}$  2040 (at 432nm) in hexane. Purify proneurosporene by chromatography on deactivated alumina [Kushwaha et al. *J Biol Chem* 245 4708 1970]. Store it in the dark, in an inert atmosphere at 0°. [*Beilstein* 1 IV 1156.]

**Retinol (vitamin A alcohol), retinoic acid (vitamin A acid), retinyl acetate (vitamin A acetate), retinal (vitamin A aldehyde) and retnyl palmitate (vitamin A palmitate).** See in "Miscellaneous", this chapter.

**Spirilloxanthin (rhodoviolascein)** [34255-08-8] M 596.9, m 216-218°, 218°,  $\lambda_{\text{max}}$  463, 493, 528 nm,  $\epsilon_{1\text{cm}}^{1\%}$  2680 (at 493nm) in petroleum ether (b 40-70°). Spirilloxanthin crystallises from CHCl<sub>3</sub>/petroleum ether, acetone/petroleum ether, \*C<sub>6</sub>H<sub>6</sub>/petroleum ether or \*C<sub>6</sub>H<sub>6</sub> as deep red spindle-like crystals. Purify it also by chromatography on a column of CaCO<sub>3</sub>/Ca(OH)<sub>2</sub> mixture, Ca(OH)<sub>2</sub> or deactivated alumina and elute it with a

\*C<sub>6</sub>H<sub>6</sub>/MeOH mixture. It gives a blue colour with SbCl<sub>3</sub> in CHCl<sub>3</sub> with  $\lambda_{\max}$  at 642nm. [Polgar et al. *Arch Biochem Biophys* **5** 243 1944, Synthesis: Surmatis & Ofner *J Org Chem* **28** 2735 1963, Karrer & Koenig *Helv Chim Acta* **23** 460 1940.] Store it in the dark in an inert atmosphere at -20°. [Beilstein **1** III 2297, **1** IV 2750.]

**Squalane** (Cosbiol, 2,6,10,15,19,23-hexamethyltetracosane, perhydrosqualene) [111-01-3] **M 422.8, m 7-38°**, **b 176°/0.05mm, 210-215°/1mm, 274°/10mm, ~350°/760mm,  $d_4^{20}$  0.80785,  $n_D^{20}$  1.416**. Purify squalane by fractional distillation *in vacuo* or evaporative distillation. It is soluble in petroleum ether, \*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O and CHCl<sub>3</sub>, slightly soluble in alcohols, Me<sub>2</sub>CO and AcOH but insoluble in H<sub>2</sub>O. Small quantities can be purified by TLC as for squalene below. It is used as a marker in GLC and HPLC. [Staudinger & Leupold *Helv Chim Acta* **15** 223 1932, Sax & Stross *Anal Chem* **29** 1700 1951, Mandai et al. *Tetrahedron Lett* **22** 763 1981, Beilstein **1** IV 593.]

**Squalene** (all-*trans*-2,6,10,15,19,23-hexamethyl-2,6,10,14,18,22-tetracosahexaene) [111-02-4] **M 410.7, m ~75°**, **b 203°/0.1mm, 213°/1mm, 285°/25mm,  $d^{25}$  0.8670,  $n_D^{20}$  1.4905**. Crystallise squalene repeatedly from Me<sub>2</sub>CO (1.4ml/g) using a Dry-ice bath, wash the crystals with cold acetone, then freeze the squalene under vacuum. Squalene is further purified by passage through a column (or TLC plate) of silica gel or chromatographed on activated alumina, using petroleum ether as eluent and stored in a vacuum in the dark. Dauben et al. [*J Am Chem Soc* **74** 4321 1952] purified squalene *via* its hexachloride which is bactericidal. It is a key intermediate in sterol biosynthesis. [Capstack et al. *J Biol Chem* **240** 3258 1965, Krishna et al. *Arch Biochem Biophys* **114** 200 1966, Heilbron & Thompson *J Chem Soc* 883 1929, Karrer et al. *Helv Chim Acta* **13** 1084 1930, UV: Farmer et al. *J Chem Soc* 544 1943, Beilstein **1** III 1071.]

**Violaxanthin** (5*S*,6*S*,5'*S*,6'*S*-diepoxyzeaxanthin) [126-29-4] **M 600.9, m 200°, 207-208°,  $[\alpha]_{C_d}^{18}$  +35°** (c 0.1, CHCl<sub>3</sub>). Crystallise violaxanthin from MeOH (orange prisms) by adding a little Et<sub>2</sub>O in MeOH, or CS<sub>2</sub> (red brown crystals). Also purify it by column chromatography, and the purity is checked by TLC (see  $\delta$ -carotene). It has  $\lambda_{\max}$  at 415, 440 and 469nm. [Kuhn & Winterstein *Ber* **64** 326 1931, Karrer et al. *Helv Chim Acta* **14** 1044 1931, Absolute Config: Bartlett et al. *J Chem Soc (C)* 2527 1969, Beilstein **19** III/IV 1139.]

**Xanthophyll** (Lutein, 3*R*,3'*R*- $\alpha$ -carotene-3,3'-diol) [127-40-2] **M 568.9, m 183°, 193°, 196°,  $[\alpha]_{C_d}^{18}$  +165°** (c 0.7, \*C<sub>6</sub>H<sub>6</sub>)[NB: Cd I and II have  $\lambda$  533 & 537 nm],  $\epsilon_{1cm}^{1\%}$  ( $\lambda_{\max}$ ) **1750 (423nm), 2560 (446nm), 2340 (477.5nm) and  $\epsilon_{mol}$  144,800 (at 445nm) in EtOH;  $\lambda_{\max}$  in CS<sub>2</sub> 446, 479 and 511nm, and  $\epsilon_{mol}$  127,000 (at 458nm, \*C<sub>6</sub>H<sub>6</sub>)**. Crystallise lutein from MeOH (1g/700ml; copper-coloured prisms) or from diethyl ether by adding MeOH. Also purify it by chromatography on columns of magnesia or calcium hydroxide, and recrystallise it from CS<sub>2</sub>/EtOH. It may be purified *via* the dipalmitate ester which crystallises from EtOH with m 92°. Store it in the dark, in an inert atmosphere, e.g. Ar. [Buchecker et al. *Helv Chim Acta* **57** 631 1974, Karrer *Helv Chim Acta* **34** 2060 1951, Absolute Config: Goodfellow et al. *J Chem Soc, Chem Commun* 1578 1970, Beilstein **6** II 1026, **6** III 5871, **6** IV 7017.]

**Zeaxanthin** [all *trans*- $\beta$ -carotene-3,3'(R,R')-diol] [144-68-3] **M 568.9, m 205-206° (natural), 207°, 215.5°,  $[\alpha]_{C_d}^{18}$  -44°** (c 0.7, CHCl<sub>3</sub>), [NB: Cd I and II have  $\lambda$  533 & 537 nm],  $\lambda_{\max}$  **275nm (log  $\epsilon$  4.34), 450nm (log  $\epsilon$  5.14), 480nm (log  $\epsilon$  5.07) in EtOH,  $\epsilon_{mol}$  132,900 (at 452nm, Me<sub>2</sub>CO)**. Zeaxanthin forms yellow plates (with a blue lustre) from MeOH, EtOH and with 0.5 MeOH from \*C<sub>6</sub>H<sub>6</sub>/MeOH. The diacetate (from natural) has m 154-155° (from CH<sub>2</sub>Cl<sub>2</sub>/MeOH), the UV has  $\lambda_{\max}$  ( $\epsilon_{1cm}^{1\%}$ ) at 452, 480nm (2210, 1945). The dipalmitate (Physalein) (from natural source) has m 95-96° (from CH<sub>2</sub>Cl<sub>2</sub>/MeOH), the UV:  $\lambda_{\max}$  ( $\epsilon_{1cm}^{1\%}$ ) at 452, 480nm (1335, 1190). [Isler et al. *Helv Chim Acta* **31** 2041 1956, Hlubeck et al. *J Chem Soc, Perkin Trans I* 848 1974.] The racemate is purified by chromatography on an alumina (Grade IV) column by elution with a gradient of petroleum ether (b 60-80°), \*C<sub>6</sub>H<sub>6</sub> and EtOAc. The red band gives a solid which recrystallises from MeOH with m 211°, and the UV has  $\lambda_{\max}$  at 439infl, 462, 589nm ( $\epsilon$  10<sup>-3</sup>: 90, 126, 110 respectively). [Loeber et al. *J Chem Soc (C)* 404 1971, Beilstein **6** II 1026, **6** III 5865, **6** IV 7017.]

## STEROIDS

This section also includes steroidal hormones, steroidal alkaloids and cardenolides.

**(-)-3- $\beta$ -Acetoxy-5-etiolic acid** [**3- $\beta$ -acetoxy-5-etiocholonic acid, androst-5-ene-17- $\beta$ -carboxylic acid**] [51424-66-9] **M 306.5, m 238-240<sup>o</sup>, 241-242<sup>o</sup>, 243-245<sup>o</sup>, 246-247<sup>o</sup>,  $[\alpha]_D^{20}$  -19.9<sup>o</sup> (c 1, Me<sub>2</sub>CO), -36<sup>o</sup> (c 1, dioxane), -33.5<sup>o</sup> (CHCl<sub>3</sub>), pK<sub>Est</sub> ~ 4.7.** The acid is purified by recrystallisation from Me<sub>2</sub>CO, Et<sub>2</sub>O/pentane, or AcOH, dried in a vacuum oven (105<sup>o</sup>/20mm) and sublimed at high vacuum. [Staunton & Eisenbram *Org Synth* **42** 4 1962, Steiger & Reichstein *Helv Chim Acta* **20** 1404 1937.]

**21-Acetoxypregnenolone (3 $\beta$ -21-acetoxypregn-5-en-20-one)** [566-78-9] **M 374.5, m 184-185<sup>o</sup>.** Crystallise 21-acetoxypregnenolone from Me<sub>2</sub>CO by allowing the solvent to evaporate in a vacuum, then dry it *in vacuo*. The crystals become opaque on standing. [Steiger & Reichstein *Helv Chim Acta* **20** 1164 1937.]

**Adrenosterone (Reichstein's G, androst-4-ene-3,11,17-trione)** [382-45-6] **M 300.4, m 214-217<sup>o</sup>, 220-224<sup>o</sup>, 224-226<sup>o</sup>,  $[\alpha]_{546}^{20}$  +364<sup>o</sup> (c 0.18, EtOH).** Dissolve adrenosterone in Me<sub>2</sub>CO, decolorise it with charcoal, filter, add H<sub>2</sub>O, Me<sub>2</sub>CO evaporate and the solid is recrystallised from aqueous EtOH. Also recrystallise it from Et<sub>2</sub>O or Et<sub>2</sub>O/pentane and dry it at 110<sup>o</sup>/0.1mm for 2 hours. It can be sublimed under high vacuum. [Reichstein *Helv Chim Acta* **20** 953, 979 1937, Mason et al. *J Biol Chem* **116** 267 1936, Beilstein **7** III 4601.]

**Aldosterone (18-aldocorticosterone)** [52-39-1] **M 360.5, m 108-112<sup>o</sup>(hydrate), 164<sup>o</sup>(anhydrous),  $[\alpha]_D^{25}$  +161<sup>o</sup> (c 1, CHCl<sub>3</sub>).** Crystallise aldosterone from aqueous acetone. It exists in solution as an equilibrium mixture of free aldehyde and its cyclic hemiacetal, favouring the hemiacetal. The *21-acetate* crystallises from Me<sub>2</sub>CO/Et<sub>2</sub>O or CH<sub>2</sub>Cl<sub>2</sub>/EtOAc and has **m 198-199<sup>o</sup>,  $[\alpha]_D^{25}$  +121.7<sup>o</sup> (c 0.7, CHCl<sub>3</sub>).** [Barton et al. *J Chem Soc Perkin Trans 1* 2243 1975, Beilstein **8** IV 3491.]

**5 $\alpha$ -Androstane (etioallocholane)** [5 $\alpha$ - 438-22-2, 24887-75-0] **M 260.5, m 50-50.5<sup>o</sup>,  $[\alpha]_D$  +2<sup>o</sup> (c 0.12, CHCl<sub>3</sub>).** Purify 5 $\alpha$ -androstane by chromatography through Al<sub>2</sub>O<sub>3</sub> (grade III) and elute with petroleum ether and recrystallise from MeOH (2x) and Me<sub>2</sub>CO/MeOH. It sublimes at 60<sup>o</sup>/high vacuum. Also dissolve it in petroleum ether, filter it through silica gel, evaporate and recrystallise it. [Steiger & Reichstein *Helv Chim Acta* **20** 817 1937, Prelog et al. *Helv Chim Acta* **27** 66 1944, IR, NMR: Halkes & Havinga *Rec Trav Chim, Pays Bas* **84** 889 1965, Allinger et al. *Tetrahedron* **27** 5073 1971, Beilstein **5** III 1110, **5** IV 1211.]

**5 $\beta$ -Androstane (etiocholane, testane)** [438-23-3] **M 260.5, m 78-79<sup>o</sup>, 78-80<sup>o</sup>,  $[\alpha]_D^{20}$  +2.03<sup>o</sup>,  $[\alpha]_{546}^{18}$  +5<sup>o</sup> (c 1, CHCl<sub>3</sub>).** Crystallise etiocholane from acetone. The method of purification for 5 $\alpha$ -androstane (above) could be used here. [Shopee *Helv Chim Acta* **27** 246, 260 1944, Butenandt & Dannerbaum *Hoppe Seyler's Z Physiol Chem* **229** 192 1934, Beilstein **5** III 1110, **5** IV 1211.]

**4-Androstene-3,17-dione (androstenedione)** [63-05-8] **M 286.4, m 170-171<sup>o</sup>, 173-174<sup>o</sup>,  $[\alpha]_D^{20}$  +196<sup>o</sup> (c 0.13, EtOH).** Crystallise the dione from hexane. It is soluble in \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub> and EtOH. It has  $\lambda_{max}$  at 239nm. It is a precursor of estrone or testosterone and has androgenic activity. [Ruzicka & Wettstein *Helv Chim Acta* **18** 980 1935, Beilstein **7** III 3636, **7** IV 2381.]

**Androsterone (Cis) (3 $\alpha$ ,5 $\alpha$ -3-hydroxyandrost-17-one)** [53-41-8] **M 290.4, m 185-185.5<sup>o</sup>,  $[\alpha]_{546}$  +118<sup>o</sup> (c 1, EtOH),  $[\alpha]_D^{20}$  +94.6<sup>o</sup> (c 0.7, EtOH).** Crystallise androsterone from Me<sub>2</sub>CO/Et<sub>2</sub>O or Me<sub>2</sub>CO and sublime it in high vacuum. The *acetate* [1164-95-0] crystallises from Et<sub>2</sub>O, Me<sub>2</sub>CO/Et<sub>2</sub>O or aqueous EtOH and sublimes in high vacuum with **m 165-166<sup>o</sup>,  $[\alpha]_D^{25}$  +87<sup>o</sup> (c 2, EtOH).** [Ruzicka *Helv Chim Acta* **17** 1389 1934, Marker *J Am Chem Soc* **57** 1755 1935, Göndös & Orr *J Chem Soc Chem Commun* 1239 1982, Beilstein **8** IV 462.]

**epi-Androsterone (Trans)** [481-29-8] **M 290.4, m 172-173<sup>o</sup>, (161-162<sup>o</sup> dl-),  $[\alpha]_{546}$  +115<sup>o</sup> (c 1, MeOH),  $[\alpha]_D^{20}$  +44.4<sup>o</sup> (c 0.27, CHCl<sub>3</sub>).** Purify *epi*-androsterone *via* the acetate, hydrolyse this and recrystallise it from CHCl<sub>3</sub>/hexane or aqueous EtOH. The *acetate* [1239-31-2] is purified by chromatography and when crystallised

from petroleum ether has **m** 103-104°,  $[\alpha]_D^{20} +68.5^\circ$  (c 1, CHCl<sub>3</sub>). The *oxime* has **m** 194-196° (from MeOH),  $[\alpha]_D^{20} +17.5^\circ$  (c 6.2, CHCl<sub>3</sub>). The **racemic ketone** is sublimed at 130°/high vacuum and after two crystallisations from methylcyclohexane it gives prisms with **m** 161-162° (which changed crystal form at 140-145°). [Ruzicka & Wettstein *Helv Chim Acta* **18** 1264 1935, Johnson et al. *J Am Chem Soc* **75** 2275 1953, **78** 6331 1956, Cordwell et al. *J Chem. Soc* 361 1953, *Beilstein* **8** IV 462.]

**Betamethasone (9 $\alpha$ -fluoro-11 $\beta$ ,17 $\alpha$ ,21-trihydroxy-16 $\beta$ -methylpregna-1,4-diene-3,20-dione)** [378-44-9] **M** 392.5, **m** 231-136°(dec), 235-237°(dec),  $[\alpha]_D^{20} +108^\circ$  (c 1, Me<sub>2</sub>CO). Betamethasone crystallises from ethyl acetate, and has  $\lambda_{\max}$  at 238nm (log  $\epsilon$  4.18) in MeOH. The *21-acetate* [287-24-6] crystallises from Me<sub>2</sub>CO/Et<sub>2</sub>O (charcoal) **m** 196-201° (205-208°) and has  $[\alpha]_D^{20} +140^\circ$  (CHCl<sub>3</sub>). [Taub et al. *J Am Chem Soc* **82** 4012 1960, Olivetto et al. *J Am Chem Soc* **80** 6688 1958, *Beilstein* **8** IV 3501.]

**Bisnorcholeic acid (pregnane-20-carboxylic acid)** [28393-20-6] **M** 332.5, **m** 214° ( $\alpha$ -form), 242° ( $\beta$ -form), 210-211° ( $\gamma$ -form), 184° ( $\delta$ -form), 181° ( $\epsilon$ -form), **pK<sub>Est</sub>** ~5.0. Bisnorcholeic acid crystallises from EtOH ( $\alpha$ -form), or acetic acid (all forms). The  $\alpha$ -form has  $[\alpha]_D^{20} -7.5^\circ$  (c 0.8, EtOH), and its *ethyl ester* has **m** 106-107°; the  $\beta$ -form has  $[\alpha]_D^{20} +23.3^\circ$  (c 0.86, EtOH), and its *ethyl ester* is an oil; the  $\gamma$ -form has  $[\alpha]_D^{20} 0^\circ$ , and its *ethyl ester* has **m** 82-83°; the  $\delta$ -form has  $[\alpha]_D^{20} +31.7^\circ$  (c 0.76, EtOH), and its *methyl ester* has **m** 99°; the  $\epsilon$ -form has  $[\alpha]_D^{20} +14.3^\circ$  (c 0.8, EtOH), and its *methyl ester* has **m** 117°. [Wieland et al. *Hoppe Seyler's Z Physiol Chem* **161** 80 1926, *Beilstein* **9** III 2650.]

**Brassicasterol [24 $\beta$ (R)-methylcholesta-5,22E-dien-3 $\beta$ -ol, 5,22-cholestadien-24 $\beta$ (R)-methyl-3 $\beta$ -ol, 24R-Ergosta-5,22E-dien-3 $\beta$ -ol]** [474-67-9] **M** 398.7, **m** 148°, 149-150°, 157-158°,  $[\alpha]_D^{20} -60^\circ, -66^\circ$ , (c 1.2, CHCl<sub>3</sub>). It is purified from the 7,22E-isomer, which is formed during synthesis, by TLC on Silica Gel G-Celite (1:1) with 30% EtOAc/hexane as eluant and recrystallised from MeOH or EtOH [Anastasia et al. *J C S Perkin I* 379 1983]. In a different synthesis from enantiomeric 23,24-cyclopropane precursor 3 $\beta$ -acetates, [22E,24 $\beta$ (R)-3 $\beta$ -brassicasteryl acetate **m** 152-154°], is obtained and purified by HPLC [using two *Altex Ultrasphere*<sup>TM</sup> ODS columns 5 $\mu$ m (10mm id x 25cm) and MeOH mobile phase], has **t<sub>R</sub>** = 118 minutes and crystallises from MeOH. The *acetate* has the <sup>1</sup>H NMR (360MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 5.37 (m, 1H, H-C6), 5.18 (m, 2H, H-C22 and H-C23), 4.60 (m, 1H, H-C3), 2.03 (s, 3H, OCOCH<sub>3</sub>), 1.02 (s, 3H, H<sub>3</sub>C19), 1.01 (d, 3H, *J* = 6.49Hz, H<sub>3</sub>C21), 0.91 (d, 3H, *J* = 6.80Hz, H<sub>3</sub>C28), 0.83 (d, 3H, *J* = 6.5Hz) and 0.83 (d, 3H, *J* = 6.5Hz, H<sub>3</sub>C26), 0.81 (d, 3H, *J* = 6.5Hz, H<sub>3</sub>C27), 0.69 (s, 3H, H<sub>3</sub>C18); and HRMS *m/z* 380.3452 (100, *M*-HOAc), C<sub>28</sub>H<sub>44</sub><sup>+</sup> requires 380.3443. [Lang & Djerassi *Helv Chim Acta* **65** 407 1982, Rubinstein et al. *Phytochemistry* **15** 195 1976.] The *acetate* (5mg) in dry Et<sub>2</sub>O is de-acetylated with LAH (4 fold) at -25°/30 minutes, excess LAH is decomposed with saturated aqueous Na<sub>2</sub>SO<sub>4</sub>, H<sub>2</sub>O is added and the free sterol is extracted into Et<sub>2</sub>O, dried (Mg<sub>2</sub>SO<sub>4</sub>), evaporated and the residue is recrystallised from MeOH to give pure (22,24R)-brassicasterol **m** 156-158°, with one peak in HPLC (as above but with an ODS-2 column) at **t<sub>R</sub>** = 55 minutes. Its <sup>1</sup>H NMR (360MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 5.35 (m, 1H, H-C6), 5.181 (m, 2H, H-C22 and H-C23), 3.53 (m, 1H<sub>ax</sub>, H-C3), 1.011 (d, 3H, *J* = 6.34Hz, H<sub>3</sub>C21), 1.010 (s, 3H, H<sub>3</sub>C19), 0.910 (d, 3H, *J* = 6.84Hz, H<sub>3</sub>C28), 0.834 (d, 3H, *J* = 6.10Hz) and 0.817 (d, 3H, *J* = 6.22Hz, H<sub>3</sub>C26), 0.815 (d, 3H, *J* = 6.22Hz, H<sub>3</sub>C27), 0.693 (s, 3H, H<sub>3</sub>C18); and the HRMS has *m/z* at 398.3545 (*M*<sup>+</sup>), C<sub>28</sub>H<sub>46</sub>O requires 398.3549. [Lang & Djerassi *Helv Chim Acta* **65** 407 1982, Rubinstein et al. *Phytochemistry* **15** 195 1976, Sheikh & Djerassi *Steroids* **26** 129 1975.]

It is a plant sterol (phytosterol) present in canola oils and rape seed but is also found in some marine algae, diatoms and shellfish (oysters). It inhibits the enzyme  $\Delta^{24}$ -sterol reductase that is involved in mammalian cholesterol biosynthesis [Fernandez et al. *Biochem J* **336** 109 2002]. Brassino sterols have plant growth regulating activities [Anastasia et al. *J C S Perkin I* 379 1983.]

**Crinosterol [24 $\alpha$ (S)-methylcholesta-5,22E-dien-3 $\beta$ -ol, 5,22-cholestadien-24 $\alpha$ (S)-methyl-3 $\beta$ -ol, 24S-Campesta-5,22E-dien-3 $\beta$ -ol]** [17472-78-5] **M** 398.7, **m** 147-148°, 152-154°  $[\alpha]_D^{20} -47.2^\circ$  (c 1.2, CHCl<sub>3</sub>). Crinosterol occurs in sponges, oysters and in diatoms. It is the 24- $\alpha$ -methyl epimer of Brassicasterol (see [474-67-9] above). They sometimes occur together in Nature and sometimes are formed together in synthesis: purification of one invariably requires the removal of the other. *Crinosterol acetate* **m** 148-150°, when isolated from a synthesis was separated from the 24- $\beta$ -methyl epimer by HPLC (see brassicasterol above), had **t<sub>R</sub>** = 112 minutes, and was recrystallised from MeOH, or *alternatively*, it is prepared by acetylation of crinosterol, then

purified by TLC (10% w/w AgNO<sub>3</sub>/silica gel and developed with pure CHCl<sub>3</sub>) to give homogeneous (by GLC) acetate with **m 157-158°** (compare with above). Crinosteryl acetate has IR (Nujol) with  $\nu_{\max}$  at 1720, 970, 955 and 795 cm<sup>-1</sup>, and the <sup>1</sup>H NMR (360MHz, CDCl<sub>3</sub>, TMS) with  $\delta$  at 5.375 (m, 1H, H-C6), 5.160 (m, 2H, H-C22 and H-C23), 4.601 (m, 1H<sub>ax</sub>, H-C3), 2.033 (s, 3H, OCOCH<sub>3</sub>), 1.019 (s, 3H, H<sub>3</sub>C19), 1.002 (d, 3H,  $J = 6.68$ Hz, H<sub>3</sub>C21), 0.909 (d, 3H,  $J = 6.78$ Hz, H<sub>3</sub>C28), 0.835 (d, 3H,  $J = 6.52$ Hz) and 0.817 (d, 3H,  $J = 6.63$ Hz, H<sub>3</sub>C26), ~0.815 (d, 3H,  $J = 6.6$ Hz, H<sub>3</sub>C27), 0.690 (s, 3H, H<sub>3</sub>C18); and HRMS  $m/z$  380.3420 (100, M-HOAc), C<sub>28</sub>H<sub>44</sub> requires 380.3443. [Lang & Djerassi *Helv Chim Acta* **65** 407 1982, Anastasia et al. *J Chem Soc, Perkin I* 379 1983, Rubinstein et al. *Phytochemistry* **15** 195 1976]. Deacetylation with LAH (see brassicasterol above) gives **crinosterol** which is purified by TLC (silica gel developed with 2% v/v EtOH/CHCl<sub>3</sub>) has **m 152-154°** (reported twice, and **147-148°** reported twice, both recrystallised from MeOH or Me<sub>2</sub>CO), and has <sup>1</sup>H NMR (360MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 5.349 (m, 1H, H-C6), 5.162 (m, 2H, H-C22 and H-C23), 3.53 (m, 1H<sub>ax</sub>, H-C3), 1.010 (s, 3H, H<sub>3</sub>C19), 1.001 (d, 3H,  $J = 6.21$ Hz, H<sub>3</sub>C21), 0.910 (d, 3H,  $J = 6.86$ Hz, H<sub>3</sub>C28), 0.836 (d, 3H,  $J = 6.45$ Hz) and 0.817 (d, 3H,  $J = 6.62$ Hz, H<sub>3</sub>C26), 0.815 (d, 3H,  $J = 6.62$ Hz, H<sub>3</sub>C27), 0.693 (s, 3H, H<sub>3</sub>C18); and HRMS  $m/z$  398.3551 (M<sup>+</sup>), C<sub>28</sub>H<sub>46</sub>O requires 398.3549. [Lang & Djerassi *Helv Chim Acta* **65** 407 1982, Anastasia et al. *J Chem Soc. Perkin I* 2365 1983, Rubinstein et al. *Phytochemistry* **13** 485 1974, Rubinstein et al. *Phytochemistry* **15** 195 1976]. Brassino steroids have plant growth regulating activities [Anastasia et al. *J C S Perkin I* 379 1983].

**Campesterol (24R-24-methylcholest-5-en-3 $\beta$ -ol)** [474-62-4] **M 400.7, m 156-159°, 157-158°, [ $\alpha$ ]<sub>D</sub><sup>24</sup> -35.1° (c 1.2, CHCl<sub>3</sub>)**. Campesterol is recrystallised twice from hexane and once from Me<sub>2</sub>CO. The *benzoyl* derivative has **m 158-160°** [ $\alpha$ ]<sub>D</sub><sup>23</sup> -8.6° (CHCl<sub>3</sub>), and the *acetyl* derivative has **m 137-138°** (EtOH) and [ $\alpha$ ]<sub>D</sub><sup>23</sup> -35.1° (c 2.9, CHCl<sub>3</sub>) [Fernholz & MacPhillamy *J Am Chem Soc* **63** 1155 1941]. [Beilstein **6** III 2680.]

**CHAPS (3-[(3-cholamidopropyl)dimethylammonio]-1-propanesulfonate betaine)** [75621-03-3 anhydrous], [313223-04-0 hydrate] **M 614.89, m 157°(dec), pK<sub>Est(1)</sub> ~<2 (acidic), pK<sub>Est(2)</sub> ~10 (basic)**. Triturate this zwitterionic detergent (~50g) with MeOH (200ml), then triturate it with Me<sub>2</sub>CO (500ml), collect it by vacuum filtration and dry it thoroughly at ~20°. Recrystallise it at 0° from absolute MeOH followed by drying *in vacuo* to constant weight. It should give one spot with R<sub>F</sub> 0.32 by TLC on silica gel G in 95% MeOH/5% NH<sub>4</sub>OH and visualised with iodine vapour; the tertiary amine precursor has R<sub>F</sub> 0.40. It has low UV absorption (1% in H<sub>2</sub>O: A280nm is 0.029 and A260nm is 0.035), and the CMC (critical micellar concentration) is 8mM. It is soluble in H<sub>2</sub>O, MeOH, Me<sub>2</sub>SO but insoluble in Me<sub>2</sub>CO, MeCN, hexane, toluene and xylene. [Hjelmeland et al. *Anal Biochem* **130** 72 1983, Matuo et al. *Methods Enzymol* **198** 155 1991.]

**CHAPSO (3-[(3-cholamidopropyl)dimethylammonio]-2-hydroxy-1-propanesulfonate betaine)** [82473-24-3] **M 630.89, m 184-196°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +19° (c 2, H<sub>2</sub>O), pK<sub>Est(1)</sub> ~<2 (acidic), pK<sub>Est(2)</sub> ~9 (basic)**. Dissolve this zwitterionic detergent (~ 60g) in 40% aqueous MeOH and stir it with the mixed-bed ion-exchange resin RG-501 x 8 (100g, Bio-Rad) until the pH of the supernatant is ~ 7. Filter off the resin and evaporate the filtrate to dryness under reduced pressure. Check for homogeneity by TLC on silica gel G in 95% MeOH/5% NH<sub>4</sub>OH and visualize the spot with iodine vapour (see CHAPS above). [Hjelmeland et al. *Anal Biochem* **130** 72 1983, Matuo et al. *Methods Enzymol* **198** 155 1991.]

**Chenodeoxycholic acid (chenodiol, 3 $\alpha$ ,7 $\alpha$ -dihydroxycholan-24-oic acid)** [474-25-9] **M 392.6, m (119°, 143°, 165-167°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +14° (c 2, EtOH), pK<sub>Est</sub> ~4.9**. This major bile acid in vertebrates (~80mg) is chromatographed on silica gel (5g) and eluted with CHCl<sub>3</sub>/EtOAc (3:2) and crystallised from EtOAc/hexane. It has IR with  $\nu_{\max}$  at 1705 cm<sup>-1</sup>(CHCl<sub>3</sub>). It also crystallises from EtOAc, EtOAc/heptane after purifying *via* the poorly soluble Na and K salt if necessary. [Kametani et al. *J Org Chem* **47** 2331 1982, Beilstein **10** IV 1604.]

**5 $\alpha$ -Cholanic acid (allocholanic acid)** [546-98-4] **M 360.6, m 170, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +22° (CHCl<sub>3</sub>), pK<sub>Est</sub> ~4.9**. Purify 5 $\alpha$ -cholanic acid by chromatography on silica gel and eluting with MeOH/EtOAc/hexane (1:5:10) and recrystallising from Et<sub>2</sub>O/hexane, MeOH or AcOH. The *methyl ester* has **m 91°** (from MeOH) and <sup>1</sup>H NMR (CDCl<sub>3</sub>) has  $\delta$  for C18 & C19 at 0.66 and 0.78 [see 5 $\beta$ -cholanic acid below, Mandava et al. *Steroids* **23** 357 1974]. [Demir et al. *Org Prepn Proced Int* **19** 197 1987, Stoll et al. *Helv Chim Acta* **18** 644 1935, Beilstein **9** III 2656, **9** IV 1992.]

**5 $\beta$ -Cholanic acid (ursocholanic acid)** [546-18-9, 25312-65-6] **M 360.6, m 164-165 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  +21.7 $^{\circ}$  (CHCl<sub>3</sub>), pK<sub>Est</sub> ~4.9.** Crystallise 5 $\beta$ -cholanic acid from EtOH, Me<sub>2</sub>CO or AcOH. The *ethyl ester* has **m 93-94 $^{\circ}$**  (from 80% EtOH), **b 273 $^{\circ}$ /12mm,  $[\alpha]_{\text{D}}^{20}$  +21 $^{\circ}$  (CHCl<sub>3</sub>)**, and the *methyl ester* has **m 86-87 $^{\circ}$**  (from MeOH) after purification through Al<sub>2</sub>O<sub>3</sub> and eluting with \*C<sub>6</sub>H<sub>6</sub>, and  **$[\alpha]_{\text{D}}^{20}$  +21.2 $^{\circ}$  (CHCl<sub>3</sub>)**. The <sup>1</sup>H NMR (CDCl<sub>3</sub>) for the C18 & C19 protons has  $\delta$  at 0.65 and 0.92 [see 5 $\alpha$ -cholanic acid above, Mandava et al. *Steroids* **23** 357 1974]. [Stoll et al. *Helv Chim Acta* **18** 644 1935, Huang-Minlon *J Am Chem Soc* **71** 330 1949, Beilstein **9** III 2656, **9** IV 1992.]

**Cholesta-3,5-diene** [747-90-0] **M 368.6, m 77-78 $^{\circ}$ , 78 $^{\circ}$ , 78-79 $^{\circ}$ , 79-80 $^{\circ}$ , 79.5-80.5 $^{\circ}$ , 80-81 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{16}$  -116 $^{\circ}$  (c 1.1, CHCl<sub>3</sub>).** This diene has been prepared in many ways but several of these are from solvolysis of epicholesteryl toluene-*p*-sulfonate using KOAc/Ac<sub>2</sub>O (38 hours at 75 $^{\circ}$ ), methanolic KOAc (reflux at 3.5 hours), LAH/Et<sub>2</sub>O (20 hours at 20 $^{\circ}$ ) in *ca* 60% yield [Schmid & Kägi *Helv Chim Acta* **35** 2194 1952], AgOAc/AcOH, KOAc/MeOH or KOAc/AcOH (reflux ~3 hours 92-99% yield), neutralised Al<sub>2</sub>O<sub>3</sub> at 20 $^{\circ}$  [Evans & Shoppee *J Chem Soc* 540 1953, Shoppee et al. *J Chem Soc* 2876 1955], KOAc/aqueous Me<sub>2</sub>CO [Becker & Wallis *J Org Chem* **20** 353 1955], and diphenylguanidinium salts accelerate solvolysis [Shoppee & Williams *J Chem Soc* 686 1955]. After the usual workup involving washing with NaHCO<sub>3</sub> and H<sub>2</sub>O, the diene is separated from other steroids and impurities by chromatography through Al<sub>2</sub>O<sub>3</sub> (various grades, including activity II-III, neutralised, basic or activated at 250-310 $^{\circ}$ ) and usually eluted with pentane, pentane/\*C<sub>6</sub>H<sub>6</sub> or Et<sub>2</sub>O/MeOH. It is then recrystallised from Me<sub>2</sub>CO, Et<sub>2</sub>O/MeOH or Et<sub>2</sub>O/EtOH, dried *in vacuo* and stored at -20 $^{\circ}$ . The UV has  $\lambda_{\text{max}}$  (log  $\epsilon$ ) at 228.5 (4.308), 235 (4.344) and 243 (4.163) nm.

**Cholesta-5,7-dien-3 $\beta$ -ol (7-dehydrocholesterol, 3 $\beta$ -hydroxy-5,7-cholestadiene, Provitamin D<sub>3</sub>)** [434-16-2] **M 384.6, m 150-151 $^{\circ}$  (anhydrous), 148-152 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -113.6 $^{\circ}$  (c 1, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{20}$  -127 $^{\circ}$  (\*C<sub>6</sub>H<sub>6</sub>).** It occurs in humans and higher animals, and has been isolated from pig skin or the horned snail (*Buccinum undatum*). It is purified by recrystallisation from MeOH, Et<sub>2</sub>O/MeOH or aqueous Et<sub>2</sub>O (plates) and holds to H<sub>2</sub>O tenaciously. It is insoluble in H<sub>2</sub>O but is soluble in most organic solvents. Store it at -20 $^{\circ}$  in the dark and absence of air which oxidises it slowly. Irradiation with UV light produces vitamin D<sub>3</sub> and like all provitamin D have  $\lambda_{\text{max}}$  at 260, 270, 281 and 293.5 nm. The *acetate* [1059-86-5] has **m 129-130 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -85.3 $^{\circ}$  (c 1.2, \*C<sub>6</sub>H<sub>6</sub>)**, crystallises from MeOH, and the *benzoate* has **m 139-140 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -53.2 $^{\circ}$  (c 1.2, CHCl<sub>3</sub>)**, and crystallises from CHCl<sub>3</sub>/Me<sub>2</sub>CO in plates. [Bernstein et al. *J Org Chem* **14** 443 1949, Beilstein **6** IV 4153].

**Cholesta-5,7-dien-3 $\alpha$ -ol,** [5765-84-0] **m 125-126 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -70.5 $^{\circ}$  (CHCl<sub>3</sub>),** crystallises from MeOH in needles, and its *acetate*, **m 114.5 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -35 $^{\circ}$  (CHCl<sub>3</sub>)**, also crystallises from MeOH in needles. [Aberhart et al. *J Org Chem* **41** 1067 1976]. [for NMR of steroids see Kirk et al. *J Chem Soc Perkin Trans 2* 1567 1990, and Anastasia et al. *Steroids* **49** 543 1987, and for ORD see Weiss et al. *Tetrahedron* **21** 3105 1965].

**Cholesta-5,24-dien-3 $\beta$ -ol (Desmosterol, 24-dehydrocholesterol, 3 $\beta$ -hydroxy-5,24-cholestadiene, natural 3 $\beta$ ,20*R*-form)** [313-04-2] **M 384.6, m 117-118 $^{\circ}$ , 119-119.5 $^{\circ}$ , 120.5-121 $^{\circ}$ , 121.5 $^{\circ}$ , 121.5-122.5 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{20}$  -38.2 $^{\circ}$  (c 1.14, CHCl<sub>3</sub>).** It occurs in chick embryos, rat skins, red algae, barnacles and *Funtmia latifolia* seeds. This is a useful intermediate for the synthesis of vitamin D metabolites [Koreeda et al. *J Org Chem* **45** 1172 1980]. The sterol in hexane can be purified by chromatography on a AgNO<sub>3</sub>/silicic acid/Super-Cel column in \*C<sub>6</sub>H<sub>6</sub>, then rechromatographed on Alumina F20, eluting with \*C<sub>6</sub>H<sub>6</sub> followed by \*C<sub>6</sub>H<sub>6</sub>/EtOAc (1:1 v/v) and recrystallised from distilled MeOH or from Me<sub>2</sub>CO/MeOH. It has IR (KBr) with  $\nu_{\text{max}}$  at 1373, 1056, 1022, 959, 950, 835 and 800 cm<sup>-1</sup>; and HRMS has *m/z* 384.3421 (M<sup>+</sup>), C<sub>27</sub>H<sub>44</sub>O requires 384.3389 [Apfel *J Org Chem* **43** 2284 1978]. The *acetate* [2665-04-5] **m 99 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{28}$  -44.8 $^{\circ}$  (c 1, CHCl<sub>3</sub>)**, can be obtained by refluxing 3-chlorocholesta-5,24-diene with AcOH/AcOK for 2.5 hours, cooling, adding H<sub>2</sub>O, extracting with Et<sub>2</sub>O, drying (MgSO<sub>4</sub>), evaporating the extract, followed by chromatography through silica gel, eluting with \*C<sub>6</sub>H<sub>6</sub>, evaporating and the white crystalline residue is recrystallised from Me<sub>2</sub>CO/MeOH. The acetate has <sup>1</sup>H NMR (60MHz, CCl<sub>4</sub>, TMS) with  $\delta$  at 0.68 (s, 3H, H<sub>3</sub>C18), 0.98 (s, 3H, H<sub>3</sub>C19), 1.00 (s, 3H, H<sub>3</sub>C21), 1.58 (s, 3H, H<sub>3</sub>C26 or H<sub>3</sub>C27), 1.65 (s, 3H, H<sub>3</sub>C27 or H<sub>3</sub>C26), 1.94 (s, 3H, 3-OCOCH<sub>3</sub>), 4.50 (m, 1H<sub>ax</sub>, H-C3), 5.00 (m, H, H-C24) and 5.35 (m, 1H, H-C5). Hydrolysis of the acetate with MeOH/KOH gave *desmosterol* with **m 119-120 $^{\circ}$ ,  $[\alpha]_{\text{D}}^{28}$  -41 $^{\circ}$  (CHCl<sub>3</sub>)**. [Ochi et al. *Steroids* **30** 795 1977]. *Desmosterol* has also been obtained from the readily available 3-hydroxy-bisnorcholenic acid in several synthetic steps and finally *via* the 3-tetrahydropyranyl ether which was hydrolysed in THF containing a few drops of concentrated HCl at 40 $^{\circ}$ /5 minutes to give the *sterol* **m 120-122 $^{\circ}$**  in 70% yield

after recrystallisation from MeOH. It has  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) with  $\delta$  at 0.41 (s,  $\text{H}_3\text{C18}$ ), 0.51 (d,  $J = 6\text{Hz}$ ,  $\text{H}_3\text{C21}$ ), 0.61 (s,  $\text{H}_3\text{C19}$ ), 1.97 and 1.02 (2s,  $\text{H}_3\text{C26}$ ,  $\text{H}_3\text{C27}$ ), 2.10 (1H<sub>ax</sub>, H-C3), 3.03 (24-H) and 3.20 (6-H). [Dasgupta et al. *J Org Chem* **39** 1658 1974.] Store it at  $-20^\circ$ .

**Cholesta-5,24-dien-3 $\alpha$ -ol** [67392-80-7] **M 384.6**, has **m 137-139 $^\circ$**  (from MeOH),  $[\alpha]_{\text{D}}^{28} -43.5^\circ$  ( $\text{CHCl}_3$ ), and its **acetate** [67383-62-4] **M 426.6**, has **m 112-115 $^\circ$**  (from MeOH),  $[\alpha]_{\text{D}}^{28} -47.5^\circ$  ( $\text{CHCl}_3$ ).

The **3 $\beta$ ,20S-form** [59532-46-6] **M 384.6**, has **m 125-127 $^\circ$**  (from MeOH), obtained by dehydration of 20S-3 $\beta$ ,25-dihydroxycholest-5-ene with dioxane/ $\text{H}_2\text{SO}_4$  (50% yield) has IR (KBr) with  $\nu_{\text{max}}$  at 3000 (OH)  $\text{cm}^{-1}$ , and  $^1\text{H}$  NMR (60MHz,  $\text{CDCl}_3$ , TMS) with  $\delta$  at 0.70 (s,  $\text{H}_3\text{C18}$ ), 0.84 (d,  $\text{H}_3\text{C21}$ ), 1.02 (s,  $\text{H}_3\text{C19}$ ), 1.62 and 1.70 (2s,  $\text{H}_3\text{C26}$ ,  $\text{H}_3\text{C27}$ ). It is an inhibitor of the cholesterol side-chain cleavage enzyme system in bovine adrenocortical preparations. [Bernstein et al. *Steroids* **27** 361 1976.]

**5 $\alpha$ -Cholestane** [481-21-0] **M 372.7**, **m 80 $^\circ$ , 80-80.5 $^\circ$** ,  $[\alpha]_{546}^{20} +29.5^\circ$ ,  $[\alpha]_{\text{D}}^{20} +24^\circ$ ,  $[\alpha]_{\text{D}}^{18} +30.2^\circ$ , (**c 2**,  $\text{CHCl}_3$ ). 5 $\alpha$ -Cholestane is prepared by catalytic reduction of cholestene  $\text{EtOAc}/\text{PtO}_2/\text{H}_2$ , and recrystallised from  $\text{EtOAc}$  or  $\text{Et}_2\text{O}/\text{EtOH}$ . Store at room temperature. The  $^1\text{H}$  NMR (100MHz,  $\text{CDCl}_3$ , TMS) signals of the methyl groups have  $\delta$  at 0.645 (s, 13-Me), 0.778 (s, 10-Me), 0.898 (,  $J = 6.0\text{Hz}$ , 20-Me) and 0.861 (d,  $J = 6.3\text{Hz}$ , 25-Me<sub>2</sub>) [Wittstruck et al. *J Chem Soc, Perkin Trans I* 1403 1977], and its CD in hexane has a  $\Delta\epsilon$  value of +0.5 at the longest wavelength of 183nm [Kirk et al. *Tetrahedron Lett* 1355 1973]. Note that catalytic hydrogenation of cholest-4-ene in acidic medium provides 5- $\alpha$ -cholestane whereas hydrogenation in neutral media results in 5- $\beta$ -cholestane [Windaus *Chem Ber* **52** 170 1919]. [Ruzicka et al. *Helv Chim Acta* **16** 327, 334 1933, *Beilstein* **5** III 1132, **5** IV 1227.]

**5 $\beta$ -Cholestane (coprostane,  $\psi$ -cholestane, pseudocholestane)** [481-20-9, 25269-18-5 + 21 other numbers + 105308-41-6] **M 372.7**, **m 70 $^\circ$ , 71-72 $^\circ$** ,  $[\alpha]_{\text{D}}^{20} +25^\circ$  (**c 2**,  $\text{CHCl}_3$ ). Catalytic hydrogenation ( $\text{Pt}_2\text{O}/\text{Et}_2\text{O}$ ) of 3-chlorocholest-4-ene followed by crystallisation from  $\text{Me}_2\text{CO}$  [Young et al. *J Am Chem Soc* **81** 1452 1959], or catalytic reduction ( $\text{Pt}_2\text{O}/\text{Et}_2\text{O}$ ) followed by chromatography through an  $\text{Al}_2\text{O}_3$  column and eluting with petroleum ether followed by recrystallisation of the residual oil from  $\text{Et}_2\text{O}/\text{EtOH}$  give high yields of coprostane [Nickson & Bagli *J Am Chem Soc* **83** 1498 1961] (see previous entry). It also recrystallises in needles from  $\text{EtOH}$ . Store it at room temperature. Its CD in hexane has a  $\Delta\epsilon$  value of +1.1 at the longest wavelength of 184nm (see previous entry) [Kirk et al. *Tetrahedron Lett* 1355 1973].

**5 $\alpha$ -Cholestan-3 $\beta$ -ol ( $\beta$ -cholestanol, dihydrocholesterol)** [80-97-7] **M 388.7**, **m 142-143 $^\circ$ (monohydrate)**,  $[\alpha]_{546}^{20} +28^\circ$  (**c 1**,  $\text{CHCl}_3$ ),  $[\alpha]_{\text{D}} +27.4^\circ$  (in  $\text{CHCl}_3$ ). Purify 5- $\alpha$ -cholestan-3 $\beta$ -ol via acetylation, crystallisation and de-acetylation, then recrystallisation from  $\text{EtOH}$  or slightly aqueous  $\text{EtOH}$ , or  $\text{MeOH}$ . Its solubility is: 0.5% ( $\text{MeOH}$ ) and 1% ( $\text{EtOH}$ ) at  $25^\circ$ . [Mizutani & Whitten *J Am Chem Soc* **107** 3621 1985.] The **acetate** has **m 114-115 $^\circ$**  from  $\text{EtOAc}/\text{MeOH}$  and  $[\alpha]_{\text{D}}^{20} +13^\circ$  (**c 2**,  $\text{CHCl}_3$ ). [Bruce & Ralls *Org Synth Col Vol II* 191 1943, *Beilstein* **6** IV 3577.]

**5 $\beta$ -Cholestan-3 $\alpha$ -ol (Epicoprostanol, epi-5 $\beta$ -cholestanol)** [516-88-7] **M 388.7**, **m 115-116 $^\circ$ , 116-117 $^\circ$ , 118 $^\circ$** ,  $[\alpha]_{\text{D}}^{20} +32^\circ$  (**c 1.8**,  $\text{CHCl}_3$ ). It is the main product from the acetolysis of 3 $\beta$ -chlorocoprostanol (KOAc/AcOH, reflux for 6 hours) involving a Walden inversion, and purified as in the following entry [Bridgewater & Shoppee *J Chem Soc* 1709 1953]. It has also been obtained by reduction of 5 $\beta$ -coprostan-3-one with LAH to give a 94% yield of epicoprostanol m 115-116 $^\circ$ , which did not form a precipitate with digitonin, and the small amount of digitonide formed gave a small quantity of coprostanol m 100 $^\circ$  (~4% yield and crystallised from  $\text{MeOH}$ , see below) [Shoppee & Summers *J Chem Soc* 687 1950]. Alternatively, the residue is purified as described in the succeeding entry or recrystallised from  $\text{Me}_2\text{CO}$  or  $\text{Et}_2\text{O}/\text{C}_6\text{H}_6$  to give the desired **3 $\alpha$ -ol**, after drying *in vacuo* at  $80^\circ$ . Note that crystals from  $\text{EtOH}$  have **m 115-116 $^\circ$** , and from  $\text{Me}_2\text{CO}$  have **m 117-118 $^\circ$** . Store it at room temperature. The **3 $\alpha$ -yl-acetate** (obtained by catalytic reduction of coprostan-3-one,  $\text{Et}_2\text{O}/\text{MeOH}/\text{PtO}_2/\text{H}_2$ , followed by chromatographic separation as below and acetylation) has **m 87-89 $^\circ$ , 89-91 $^\circ$** ,  $[\alpha]_{\text{D}}^{18} +42^\circ$  (**c 1.42**,  $\text{CHCl}_3$ ). [Shoppee & Summers *J Chem Soc* 687 1950, Ruzicka et al. *Helv Chim Acta* **17** 1407 1934.]

**5 $\beta$ -Cholestan-3 $\beta$ -ol (coprostan-3 $\beta$ -ol)** [360-68-9] **M 388.7**, **m 100-101 $^\circ$ , 101 $^\circ$** ,  $[\alpha]_{\text{D}}^{20} +28^\circ$  (**c 1.8**,  $\text{CHCl}_3$ ). It is the main product from the acetolysis of 3 $\alpha$ -chlorocoprostanol (KOAc/AcOH, reflux for 6 hours) involving a Walden inversion. Unlike epicoprostanol (above), coprostanol is precipitated by a 1.33% solution of digitonin in

95% EtOH, the digitonide can be collected, decomposed by dissolving in pyridine, precipitating the digitonin with excess of Et<sub>2</sub>O, filtering off digitonin and the coprostanol is obtained by evaporation of the organic solvent. A possible impurity is the 3 $\alpha$ -ol from which it can be separated by chromatography on Al<sub>2</sub>O<sub>3</sub> and eluting with Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> (1:24) whereby the 3 $\alpha$ -ol runs first (**m** 116-117°, see preceding entry) and further elution gives the pure 3 $\beta$ -coprosterol (**m** 99-101°). Crystallise it from MeOH (needles) or aqueous EtOH. Store it at room temperature. The 3 $\beta$ -yl-acetate from acetylation with Ac<sub>2</sub>O (also obtained by catalytic reduction of coprostan-3-one, AcOH/PtO<sub>2</sub>/H<sub>2</sub>, followed by chromatographic separation as above and acetylation) has **m** 91°, **90-91°**, [ $\alpha$ ]<sub>D</sub><sup>18</sup> +14.5° (**c** 1.8, CHCl<sub>3</sub>). Its solubility is 0.79% in H<sub>2</sub>O at 25°. [Bridgewater & Shoppee *J Chem Soc* 1709 1953, Shoppee & Summers *J Chem Soc* 687 1950, see Ruzicka et al. *Helv Chim Acta* 17 1407 1934.]

**Cholest-2-ene** [15910-23-3] **M** 370.6, **m** 75-76°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +66° (**c** 1.65, CHCl<sub>3</sub>). Purify cholest-2-ene by chromatography on Al<sub>2</sub>O<sub>3</sub> (Spence H) and elute with petroleum ether (b 40-60°), and recrystallise the residue from EtOAc/MeOH. Also recrystallise it from MeOH or diethyl ether/acetone. [Alt & Barton *J Chem Soc* 4284 1954, Bergbreiter & Chandran *J Am Chem Soc* 109 174 1987, *Beilstein* 5 III 1320, 5 IV 1507.]

**Cholesterol (cholest-5-en-3 $\beta$ -ol)** [57-88-5] **M** 386.7, **m** 148.9-149.4°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -35° (hexane). Crystallise cholesterol from ethyl acetate, EtOH or isopropyl ether/MeOH. [Hiromitsu & Kevan *J Am Chem Soc* 109 4501 1987.] For extensive details of purification through the dibromide, see Fieser [*J Am Chem Soc* 75 5421 1953] and Schwenk and Werthessen [*Arch Biochem Biophys* 40 334 1952], and by repeated crystallisation from acetic acid; see Fieser [*J Am Chem Soc* 75 4395 1953]. Like many sterols, cholesterol gives colour reactions with conc H<sub>2</sub>SO<sub>4</sub>: When cholesterol is dissolved in a small volume of CHCl<sub>3</sub> and mixed with conc H<sub>2</sub>SO<sub>4</sub>, the colour of the organic layer becomes crimson, then changes to purple and on further standing in air it turns to blue, then green and finally yellow. The H<sub>2</sub>SO<sub>4</sub> layer develops a green fluorescence. [*Beilstein* 6 III 2607, 6 IV 4000.]

**Cholesteryl acetate** [604-35-3] **M** 428.7, **m** 113-115°, 115-116°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -51° (**c** 5, CHCl<sub>3</sub>). Crystallise the acetate from *n*-pentanol or Me<sub>2</sub>CO. Also purify it by chromatography through silica gel and eluting with MeOH. [*Beilstein* 6 III 2607, 6 IV 4004.]

**Cholesteryl myristate** [1989-52-2] **M** 597.0, **m** 69-71°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -25.4° (**c** 1, CHCl<sub>3</sub>). Crystallise the myristate ester from Me<sub>2</sub>CO, EtOH/EtOAc or EtOH/Et<sub>2</sub>O or *n*-pentanol. Purify it also by column chromatography on silica gel and eluting with MeOH then evaporating to dryness. Recrystallise it and finally, dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub> and store it at -20°. [Labarière et al. *Analyt Chem* 30 1466 1958, Malanik & Malat *Anal Chim Acta* 76 464 1975, *Beilstein* 6 III 2638].

**Cholesteryl oleate** [303-43-5] **M** 651.1, **m** 48.8-49.4°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -24° (**c** 1, CHCl<sub>3</sub>). Purify the oleate ester by chromatography on silica gel and eluting with MeOH. [*Beilstein* 6 III 2642.]

**3 $\alpha$ -5 $\beta$ -7 $\alpha$ -12 $\alpha$ -Cholic acid (3 $\alpha$ -7 $\alpha$ -12 $\alpha$ -trihydroxy-5 $\beta$ -cholan-24-oic acid, Cholalin)** [81-25-4] **M** 408.6 (anhydrous), 426.6 (hydrate) **m** 196-198°, 198-200°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +37°, [ $\alpha$ ]<sub>D</sub><sup>546</sup> +41° (**c** 0.6, EtOH), pK<sup>20</sup> 4.98. This bile acid crystallises from H<sub>2</sub>O, aqueous AcOH (as hydrate), wet Et<sub>2</sub>O (as hydrate) or EtOH (as alcoholate). Dry it under vacuum at 94° to give the *anhydrous acid* **m** 198°. Its solubility (w/v) at 15° is 0.028% in H<sub>2</sub>O, 3.06% in EtOH, 0.12% in Et<sub>2</sub>O, 0.51% in CHCl<sub>3</sub>, 0.036% in \*C<sub>6</sub>H<sub>6</sub>, 2.82% in Me<sub>2</sub>CO and 15.2% in AcOH. When an alcoholic solution of cholic acid + I<sub>2</sub> is added to aqueous KI, it forms a molecular compound (C<sub>24</sub>H<sub>40</sub>O<sub>5</sub>I)<sub>4</sub> KI. H<sub>2</sub>O. The *methyl ester* is dimorphic with **m** 155-156° and 162° (from 95% aqueous EtOH), and [ $\alpha$ ]<sub>D</sub><sup>20</sup> +25° (EtOH), and the *ethyl ester* has **m** 155° (162-163° after recrystallisation from EtOAc/petroleum ether b 30-60° 2:8, Cortese *J Am Chem Soc* 59 2532 1937). [Anderson et al. *Biochem J* 67 323 1957, 85 236 1962, Ekwall et al. *Acta Chem Scand* 11 50 1957.] The acid has been used as a non-denaturing ionic detergent, and for solubilising membrane-bound proteins in their native state [Hooper *Biochem Soc Trans* 14 586 1986]. It has also been used in liposome preparations [Yang & Lundahl *Anal Biochem* 218 210 1994.] [*Beilstein* 10 III 2162, 10 IV 2072.]

**4,5-Coprosten-3-ol (cholest-4-ene-3 $\beta$ -ol, allocholesterol)** [517-10-2] **M** 386.7, **m** 132°, [ $\alpha$ ]<sub>D</sub><sup>24</sup> +43.7°, (**c** 1, \*C<sub>6</sub>H<sub>6</sub>). Purify 4,5-coprosten-3-ol by dissolving it in Et<sub>2</sub>O, adding an equal volume of MeOH and removing the

Et<sub>2</sub>O with a stream of CO<sub>2</sub> until crystallisation begins. The sterol crystallises in needles when cooled in an ice-salt bath. Dry it *in vacuo*. The *acetate* crystallises from aqueous MeOH **m** 85°. [Schoenheimer & Evans *J Biol Chem* **114** 567 1936, Stoll *Hoppe Seyler's Z Physiol Chem* **246** 10 1937, *Beilstein* **6** IV 3577.]

**Corticosterone (11β, 21-dihydroxypregn-4-en-3,20-dione)** [50-22-6] **M 346.5, m 180-181°, 180-182°, 181-184°, [α]<sub>D</sub><sup>25</sup> +223° (c 1.1, EtOH), [α]<sub>D</sub><sup>23-25</sup> +194° (c 0.1, dioxane)**. Purify corticosterone by recrystallisation from Me<sub>2</sub>CO (trigonal plates), EtOH or isoPrOH. It has UV with λ<sub>max</sub> at 240nm, and gives an orange-yellow solution with strong fluorescence on treatment with concentrated H<sub>2</sub>SO<sub>4</sub>. It is insoluble in H<sub>2</sub>O but soluble in organic solvents. [Reichstein & Euw *Helv Chim Acta* **21** 1197 1938, **27** 1287 1944; Mason et al. *J Biol Chem* **114** 613 1936; ORD: Foltz et al. *J Am Chem Soc* **77** 4359 1955; NMR: Shoolery & Rogers *J Am Chem Soc* **80** 5121 1958.] The 21-*O*-*acetyl* derivative [1173-26-8] crystallises from Me<sub>2</sub>CO/Et<sub>2</sub>O with **m** 152.5-153°, [α]<sub>D</sub><sup>20</sup> +195° (c 0.6, Me<sub>2</sub>CO), and the 21-*O*-*benzoyl* derivative crystallises from AcOH/Et<sub>2</sub>O with **m** 201-202° [Reichstein *Helv Chim Acta* **20** 953 1937]. [*Beilstein* **8** IV 2907.]

**Cortisol** See hydrocortisone [50-23-7].

**Cortisone** [53-06-5] **M 360.5, m 230-231°, [α]<sub>D</sub><sup>20</sup> +225° (c 1, EtOH)**. Crystallise cortisone from 95% EtOH or acetone. The UV has ε 14,000 M<sup>-1</sup>cm<sup>-1</sup> at 237nm (EtOH). [Hems *J Pharm Pharmacol* **5** 409 1953, *Beilstein* **8** IV 3480.]

**Cortisone-21-acetate** [50-04-4] **M 402.5, m 242-243°, [α]<sub>D</sub><sup>20</sup> +227° (c 1, CHCl<sub>3</sub>)**. Crystallise cortisone-21-acetate from acetone or CHCl<sub>3</sub>. The UV has ε 15,800 M<sup>-1</sup>cm<sup>-1</sup> at 238nm in dioxane. [Sarett *J Biol Chem* **162** 601 1946, *Beilstein* **8** III 4058, **5** IV 3481.]

**Deoxycholic acid** [83-44-3] **M 392.6, m 171-174°, 176°, 176-178°, [α]<sub>D</sub><sup>20</sup> +64° (c 1, EtOH), [α]<sub>D</sub><sup>20</sup> +55° (c 2.5, EtOH), pK<sup>25</sup> 6.58**. Reflux the acid with CCl<sub>4</sub> (50ml/g), filter, evaporate under vacuum at 25°, recrystallise the residue from acetone and dry it under vacuum at 155° [Trenner et al. *J Am Chem Soc* **76** 1196 1954]. A solution of (cholic acid-free) material (100ml) in 500ml of hot EtOH is filtered, evaporate it to less than 500ml on a hot plate, and pour it into 1500ml of cold diethyl ether. The precipitate, filtered off by suction, is crystallised twice from 1-2 parts of absolute EtOH, to give an *alcoholate*, **m** 118-120°, which is dissolved in EtOH (100ml for 60g) and poured into boiling water. After boiling until free of the EtOH, the precipitate is filtered off, dried, ground and dried to constant weight *in vacuo* [Sobotka & Goldberg *Biochem J* **26** 555 1932]. Deoxycholic acid is also freed from fatty acids and cholic acid by silica gel chromatography and elution with 0.5% acetic acid in ethyl acetate [Tang et al. *J Am Chem Soc* **107** 4058 1985]. It can also be recrystallised from butanone. Its solubility in H<sub>2</sub>O at 15° is 0.24g/L, but in EtOH it is 22.07g/L. [*Beilstein* **10** IV 1608.]

**11-Deoxycorticosterone (21-hydroxyprogesterone)** [64-85-7] **M 330.5, m 141-142°, [α]<sub>D</sub><sup>20</sup> +178° and [α]<sub>D</sub><sup>546</sup> +223° (c 1, EtOH)**. Crystallise 11-deoxycorticosterone from diethyl ether. [Schindler et al. *Helv Chim Acta* **24** 360 1941, Steiger & Reichstein *Helv Chim Acta* **20** 1164 1937.]

**11-Deoxycorticosterone acetate (21-acetoxy-4-pregnen-3,20-dione)** [56-47-3] **M 372.5, m 154-159°, 154-160°, 155-157°, 155-161°, [α]<sub>D</sub><sup>20</sup> +174° (c 1, dioxane), [α]<sub>D</sub><sup>22-24</sup> +196° (c 1, CHCl<sub>3</sub>)**. 11-Deoxycorticosterone acetate recrystallises from EtOH as needles or Me<sub>2</sub>CO/hexane, and sublimes at high vacuum. It is partly soluble in MeOH, Me<sub>2</sub>CO, Et<sub>2</sub>O and dioxane but insoluble in H<sub>2</sub>O. [Reichstein & Euw *Helv Chim Acta* **23** 136 1940, Romo et al. *J Am Chem Soc* **79** 5034 1957; NMR: Shoolery & Rogers *J Am Chem Soc* **80** 5121 1959, *Beilstein* **8** IV 2195.]

**Dexamethasone (9-α-fluoro-16-α-methylprednisolone, prednisolone F)** [50-02-2] **M 392.5, m 262-264°, 268-271°, [α]<sub>D</sub><sup>25</sup> +77.5° (c 1, dioxane)**. Dexamethasone has been recrystallised from Et<sub>2</sub>O or small volumes of EtOAc. Its solubility in H<sub>2</sub>O is 10 mg/100ml at 25°; and is freely soluble in Me<sub>2</sub>CO, EtOH and CHCl<sub>3</sub>. Store it below 8°. It is a glucocorticoid steroidal anti-inflammatory agent. A stock solution of 20μg/ml is generally made by adding 1ml of absolute EtOH to 100μg of steroid, stir gently to dissolve then add 49ml of sterile medium while mixing. [Arth et al. *J Am Chem Soc* **80** 3161 1958; for the β-methyl isomer see Taub et al. *J Am*

*Chem Soc* **82** 4025 1960, see *Beilstein* **8** IV 3501.]

**Dexamethasone 21-acetate** (9- $\alpha$ -fluoro-16- $\alpha$ -methylprednisolone-21-acetate, prednisolone F acetate) [1177-87-3] **M 434.5, m 215-225<sup>o</sup>, 229-231<sup>o</sup>,  $[\alpha]_D^{25} +77.6^{\circ}$  (c 1, dioxane), +73<sup>o</sup> (c 1, CHCl<sub>3</sub>).** Dexamethasone 21-acetate is purified on neutral Al<sub>2</sub>O<sub>3</sub> using CHCl<sub>3</sub> as eluent, the fractions are evaporated, and the residue is recrystallised from CHCl<sub>3</sub>. Store it below 8<sup>o</sup>. It has  $\lambda_{\max}$  at 239nm ( $\epsilon$  14,900) in EtOH. It is a glucocorticoid steroidal anti-inflammatory agent. [Oliveto et al. *J Am Chem Soc* **80** 4431 1958]. [*Beilstein* **8** IV 3501.]

**Dexamethasone 21-phosphate disodium salt** (Baldex, Colvasone, Dexabene, Desone among many trade names) [2392-39-4] **M 516.4, m 233-255<sup>o</sup>,  $[\alpha]_D^{25} +74^{\circ}$  (for anhydrous salt, c 1, H<sub>2</sub>O).** Recrystallise it from EtOH and dry it *in vacuo*. It has  $\lambda_{\max}$  at 238-239nm ( $\epsilon$  14,000) in EtOH. Store it below 8<sup>o</sup>, but ethanolic solutions should be stored at ~0<sup>o</sup>.

Like the preceding two drugs it is also an anti-inflammatory drug, but is more soluble in H<sub>2</sub>O. It is a pro-drug and is converted to dexamethasone *in vivo*; it also stimulates glutamine uptake in the cerebral cortex.

**Digitoxigenin** (3 $\beta$ ,14 $\beta$ -dihydroxy-20(22)-norcholenic acid lactone) [143-62-4] **M 374.5, m 253<sup>o</sup>, 249-255<sup>o</sup>,  $[\alpha]_{546}^{20} +17^{\circ}$  (c 1, MeOH).** Dissolve digitoxigenin in EtOAc, wash it with H<sub>2</sub>O, dry it (Na<sub>2</sub>SO<sub>4</sub>), evaporate and recrystallise it from 40% aqueous EtOH or aqueous MeOH then EtOAc/Et<sub>2</sub>O. The UV has  $\lambda_{\max}$  at 218nm ( $\epsilon$  14,500 M<sup>-1</sup>cm<sup>-1</sup>). The 3-acetate crystallises from Me<sub>2</sub>CO/Et<sub>2</sub>O with **m 222-227<sup>o</sup>,  $[\alpha]_D^{20} +21^{\circ}$  (c 1, CHCl<sub>3</sub>).** [Wyss et al. *Helv Chim Acta* **43** 644 1960, Cruz et al. *J Org Chem* **42** 3580 1977, *Beilstein* **18** IV 1468.]

**Dihydrotachysterol** [67-96-9] **M 398.7, m 125-127<sup>o</sup>,  $[\alpha]_D^{20} +97^{\circ}$  (CHCl<sub>3</sub>).** Crystallise the sterol from 90% MeOH, and UV with  $\lambda_{\max}$  at 242, 251 and 261nm ( $E_{1\text{cm}}^{1\%}$  760, 1010 and 650) in EtOH. The acetate has **m 108-110<sup>o</sup>** and  $[\alpha]_D^{20} +32.8^{\circ}$  (CHCl<sub>3</sub>), and UV with  $\lambda_{\max}$  at 242, 251 and 261nm ( $E_{1\text{cm}}^{1\%}$  780, 910 and 600) in EtOH. The propionate has **m 97-98<sup>o</sup>** and  $[\alpha]_D^{20} +37^{\circ}$  (CHCl<sub>3</sub>), and UV with  $\lambda_{\max}$  242, 251 and 261nm ( $E_{1\text{cm}}^{1\%}$  750, 860 and 570) in EtOH. [Werder Hoppe Seyler's *Z Physiol Chem* **260** 119 1939, Windaus et al. *Justus Liebigs Ann Chem* **499** 1978 1932, *Beilstein* **6** III 2833, **6** IV 3994, 4161.]

**Diosgenin (25R-spirostaen-3 $\beta$ -ol)** [512-04-9] **M 294.5, m 204-207<sup>o</sup>,  $[\alpha]_D^{25} -129^{\circ}$  (in Me<sub>2</sub>CO).** Crystallise diosgenin from acetone, and chromatograph it on Al<sub>2</sub>O<sub>3</sub> and elute with \*C<sub>6</sub>H<sub>6</sub>/Et<sub>2</sub>O (9:1), then recrystallise it from MeOH. Its solubility is ~4% in H<sub>2</sub>O and 5% in CHCl<sub>3</sub>. The acetate crystallises from AcOH with **m 198<sup>o</sup>**; and has  $[\alpha]_D^{20} -119^{\circ}$  (pyridine). [Marker et al. *J Am Chem Soc* **65** 1199 1943. Mazur et al. *J Am Chem Soc* **82** 5889 1960, *Beilstein* **19** IV 862.]

**$\alpha$ -Ecdyson** [3604-87-3] **M 464.7, m 239-242<sup>o</sup>, 242<sup>o</sup>,  $[\alpha]_D^{20} +72^{\circ}$  (c 1, EtOH).** Recrystallise  $\alpha$ -ecdyson from tetrahydrofuran/petroleum ether and from H<sub>2</sub>O as a hydrate. It has been purified by chromatography on Al<sub>2</sub>O<sub>3</sub> and elution with EtOAc/MeOH. It has  $\lambda_{\max}$  at 242nm ( $\epsilon$  12,400). Its acetate has **m 214-216<sup>o</sup>** from EtOAc/petroleum ether, and the 2,4-dinitrophenylhydrazone has **m 170-175<sup>o</sup>(dec)** from EtOAc. [Karlson & Hoffmeister *Justus Liebigs Ann Chem* **662** 1 1963, Karlson *Pure Appl Chem* **14** 75 1967, *Beilstein* **8** IV 3613.]

**$\beta$ -Ecdyson** ( $\beta$ -echdysterone, crustecdyson, 20-hydroxyecdyson) [5289-74-7] **M 480.7, m 240-242<sup>o</sup>, 245-247<sup>o</sup>,  $[\alpha]_D^{20} +66^{\circ}$  (c 1, MeOH).** Crystallise  $\beta$ -ecdyson from water, tetrahydrofuran/petroleum ether, MeOH and EtOAc after chromatographic purification. It has  $\lambda_{\max}$  (EtOH) at 240nm ( $\epsilon$  12,670 M<sup>-1</sup>cm<sup>-1</sup>). [Also IR & NMR: Hüppi & Siddall *J Am Chem Soc* **89** 6799 1967, Kametani et al. *Tetrahedron Lett* **21** 4855 1980, *Beilstein* **8** IV 343.]

**(+)-Equilenin** [517-09-9] **M 266.3, m 258-259<sup>o</sup> (vac),  $[\alpha]_D^{16} +87^{\circ}$  (c 0.1, dioxane).** Crystallise (+)-equilenin from EtOH (solubility is 0.63% at 18<sup>o</sup>, 2.5% at 78<sup>o</sup>), aqueous EtOH or \*C<sub>6</sub>H<sub>6</sub> (Norite) and dry it in a vacuum. It sublimes on melting and at 170-180<sup>o</sup>/0.01mm. The acetate crystallises from MeOH with **m 165-167<sup>o</sup>** and  $[\alpha]_D^{20} +71^{\circ}$  (c 1, CHCl<sub>3</sub>). [Bachmann et al. *J Am Chem Soc* **62** 824 1940, *Beilstein* **8** III 1522, 1523, 1525, **8** IV 1420.]

**Ergosterol (Provitamin D<sub>2</sub>)** [57-87-4] M 396.7, m 165-166°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> -171° (c 1, CHCl<sub>3</sub>), [ $\alpha$ ]<sub>D</sub><sup>20</sup> -135° (c 1, CHCl<sub>3</sub>). Crystallise ergosterol from EtOAc, then from ethylene dichloride or EtOH/\*C<sub>6</sub>H<sub>6</sub> (3:1). It has been purified by conversion to the *isobutyl ester* which crystallises from Et<sub>2</sub>O/Me<sub>2</sub>CO (1:3) with m: turbid at 148°, melts at 159° and becomes clear at 162°, followed by hydrolysis, [Bill & Honeywell *J Biol Chem* **80** 15 1938]. When crystallised from EtOH, it forms the *1,5-hydrate* m 168°. The water is difficult to remove giving an amorphous solid m 166-183°, b 250°/high vacuum. It is light sensitive. The *benzoate* has m 169-171°, after crystallisation from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub> (4:1) after prolonged standing at 0° and becomes highly charged, with [ $\alpha$ ]<sub>D</sub><sup>20</sup> -177° (c 1, CHCl<sub>3</sub>). [UV of sterols: Hogness et al. *J Biol Chem* **120** 239 1937, *Beilstein* **6** IV 4407.]

**17 $\alpha$ -Estradiol** [57-91-0] M 272.3, m 223, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +57° (c 1, EtOH). 17 $\alpha$ -Estradiol recrystallises from aqueous EtOH (80%) as the *hemihydrate* and differs from the  $\beta$ -anomer (below) by not precipitating with digitonin in 80% aqueous EtOH. The *diacetate* [1474-52-8] crystallises from aqueous EtOH in needles with m 139-140°. The *3-benzoate* crystallises in three forms m 158°, 153° and 63°. [See references for 17 $\beta$ -Estradiol below.]

**17 $\beta$ -Estradiol (1,3,5-estratrien-3,17 $\beta$ -diol, 17- $\beta$ -estradiol was incorrectly called  $\alpha$ -estradiol)** [50-28-2] M 272.4, m 173-179°, 176-178°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +76° to +83° (c 1, dioxane). 17 $\beta$ -Estradiol (previously known as  $\alpha$ -estradiol) is purified by chromatography on SiO<sub>2</sub> (toluene/EtOAc 4:1) and recrystallised from CHCl<sub>3</sub>/hexane or 80% EtOH. It is stable in air, is insoluble in H<sub>2</sub>O, and is precipitated by digitonin. The UV has  $\lambda_{\max}$  at 225 and 280 nm. The *diacetate* [3434-88-6] has m 97-98° and forms leaflets from aqueous EtOH. The *3-benzoate* [50-50-0] crystallises from aqueous MeOH with m 193° and [ $\alpha$ ]<sub>D</sub><sup>25</sup> +58° to 63° (c 1, dioxane). [Meischer & Scholz *Helv Chim Acta* **20** 263, 1237 1937, *Biochem J* **32** 1273 1938, Oppolzer & Roberts *Helv Chim Acta* **63** 1703 1980, Inhoffen & Zühlendorf *Chem Ber* **74** 1914 1941, *Beilstein* **6** IV 6611.]

**$\beta$ -Estradiol-6-one (1,3,5-estratriene-3,17 $\beta$ -diol-6-one)** [571-92-6] M 359.4, m 278-280°, 281-283°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +4.2° (c 0.7, EtOH).  $\beta$ -Estradiol-6-one forms plates from EtOH. The *3,17-diacetate* has m 173-175° after recrystallisation from aqueous EtOH. [Longwell & Wintersteiner *J Biol Chem* **133** 219 1940.] The UV has  $\lambda_{\max}$  at 255 and 326nm in EtOH [Slaunwhite et al. *J Biol Chem* **191** 627 1951]. [*Beilstein* **8** IV 2398.]

**Estriol (1,3,5-estratrien-3 $\alpha$ ,16 $\alpha$ ,17 $\beta$ -triol)** [50-27-1] M 288.4, m 278.5-284°, 283°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +66° (c 1, dioxane). Crystallise estriol from EtOH/ethyl acetate. Also purify it by countercurrent distribution with cyclohexane/EtOAc (1:1) and EtOH/H<sub>2</sub>O (1:1). The UV (EtOH) has  $\lambda_{\max}$  at 280nm ( $\epsilon$  2,090 M<sup>-1</sup>cm<sup>-1</sup>). [Huffmann & Lott **71** 719 1949, Leeds et al. *J Am Chem Soc* **76** 2943 1954, *Beilstein* **6** IV 7550.]

**Estrone (1,3,5-Estratrien-3 $\beta$ -ol-17-one, Folliculin)** [53-16-7] M 270.4, m 260-261°, polymorphs have m 254° and 256°, [ $\alpha$ ]<sub>546</sub><sup>20</sup> +198° (c 1, dioxane), pK<sup>25</sup> 0.91. Purify estrone by chromatography on silica gel, eluting with 2:1 hexane/EtOAc and recrystallising from EtOH or Et<sub>2</sub>O/EtOH. [Danishefsky & Cain *J Am Chem Soc* **98** 4975 1976.] The *acetate* [901-93-9] crystallises from EtOH with m 125-127°. [*Beilstein* **8** III 1171.]

**Estrone 3-O-sulfamate** [148672-09-7] M 349.5. Estrone 3-O-sulfamate is purified by silica gel flash chromatography to give a product with one spot on TLC. It is an active site-directed inhibitor of estrone sulfatase, i.e. a time-dependent enzyme inactivator. [Woo et al. *J Med Chem* **39** 1349 1996, Purohit et al. *Biochemistry* **34** 11508 1995.]

**17- $\alpha$ -Ethinylestradiol** [57-63-6] M 296.4, m 141-146°, 145-146°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +4° (c 1, CHCl<sub>3</sub>). 17- $\alpha$ -Ethinylestradiol forms a *hemihydrate* on recrystallising from MeOH/H<sub>2</sub>O. It dehydrates on melting and re-melts on further heating at m 182-184°. The UV has  $\lambda_{\max}$  at 281nm ( $\epsilon$  2040) in EtOH. Its solubility is 17% in EtOH, 25% in Et<sub>2</sub>O, 20% in Me<sub>2</sub>CO, 25% in dioxane and 5% in CHCl<sub>3</sub>. [Petit & Muller *Bull Soc Chim Fr* 121 1951.] The *diacetyl* derivative has m 143-144° (from MeOH) and [ $\alpha$ ]<sub>D</sub><sup>20</sup> +1° (c 1, CHCl<sub>3</sub>) [Mills et al. *J Am Chem Soc* **80** 6118 1958]. [*Beilstein* **6** IV 6877.]

**Etiocolanic acid** [438-08-4] M 304.5, m 228-229°, pK<sub>Est</sub> ~4.7. Crystallise etiocholanic acid from glacial acetic acid and sublime it at 160°/0.002mm. The *ethyl ester* has m 77.5-78.5° (from EtOH) [ $\alpha$ ]<sub>546</sub><sup>20</sup> +52.5° (c 0.61, pyridine). [Jacobs & Elderfield *J Biol Chem* **108** 497 1935, Weiland et al. *Hoppe Seyler's Z Physiol Chem* **161** 80 1926, *Beilstein* **9** III 2644.]

**Gitoxigenin (3 $\beta$ ,14,16 $\beta$ ,21-tetrahydroxy-20(22)-norcholenic acid lactone)** [545-26-6] **M 390.5, m 223-226 $^{\circ}$ , 234 $^{\circ}$ , 239-240 $^{\circ}$  (anhydrous by drying at 60 $^{\circ}$ ),**  $[\alpha]_{\text{D}}^{20} +30^{\circ}$  (c 1, MeOH). Recrystallisation of gitoxigenin from aqueous EtOH produces plates of the *sesquihydrate* which dehydrate on drying at 100 $^{\circ}$  *in vacuo*. It also recrystallises from Me<sub>2</sub>CO/MeOH and from EtOAc (the crystals contain 1 mol of EtOAc) with  $[\alpha]_{\text{D}}^{21} +24.8^{\circ}$  (c 1, dioxane). It has UV with  $\lambda_{\text{max}}$  at 310, 485 and 520nm in 96% H<sub>2</sub>SO<sub>4</sub>. On heating with ethanolic HCl it yields *digitaligenin* with loss of H<sub>2</sub>O. [Smith *J Chem Soc* 23 1931, *Beilstein* 8 IV 2456.]

**Glycocholic acid (N-cholylglycine)** [475-31-0] **M 465.6, m 130 $^{\circ}$  (hydrate) 154-155 $^{\circ}$ , 165-168 $^{\circ}$ (anhydrous),**  $[\alpha]_{\text{D}}^{20} +37^{\circ}$  (c 1, EtOH), **pK<sup>25</sup> 4.4.** Glycocholic acid crystallises from hot water as the *sesquihydrate*. Dry it at 110 $^{\circ}$  *in vacuo*. An analytical sample is prepared by suspending the acid (4g) in H<sub>2</sub>O (400ml) at ~20 $^{\circ}$ , heating to boiling with slow stirring, filtering hot and allowing to cool to ~20 $^{\circ}$ . The acid is filtered off, washed with H<sub>2</sub>O, dried in air, recrystallised from 5% aqueous EtOH, washed well and dried over P<sub>2</sub>O<sub>5</sub> in a moderate vacuum to constant weight. Recrystallisation from EtOH/EtOAc, and drying, gave the *anhydrous* acid. [Cortese & Bauman *J Am Chem Soc* 57 1393 1935, Bergstrom & Norman *Acta Chem Scand* 7 1126 1953, *Beilstein* 10 IV 2077.]

**Glycodeoxycholic acid monohydrate (N-[3 $\alpha$ -12 $\alpha$ -dihydroxy-5 $\beta$ -cholan-24-oyl]glycine)** [360-65-6] **M 467.6, m 186-177 $^{\circ}$ (dec), 187-188 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{23} +45.9^{\circ}$  (c 1, EtOH), **pK<sub>Est</sub> ~4.4.** Glycodeoxycholic acid recrystallises from H<sub>2</sub>O or aqueous EtOH with 1 mol of H<sub>2</sub>O and is dried at 100 $^{\circ}$  *in vacuo*. Its solubility in EtOH is ~5%. [UV: Lindstedt & Sjövall *Acta Chem Scand* 11 421 1957.] The *Na* salt recrystallises from EtOH/Et<sub>2</sub>O with **m 245-250 $^{\circ}$**  and  $[\alpha]_{\text{D}}^{23} +41.2^{\circ}$  (c 1, H<sub>2</sub>O) [Wieland *Hoppe Seyler's Z Physiol Chem* 106 181 1919, Cortese *J Am Chem Soc* 59 2532 1937]. [*Beilstein* 10 IV 1611.]

**Hecogenin (25R-5 $\alpha$ -spirostan-3 $\beta$ -ol-12-one)** [467-55-0] **M 430.6, m 245-250 $^{\circ}$ , 253 $^{\circ}$ , 264-266 $^{\circ}$ , 268 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{23} +8^{\circ}$  (c 1, CHCl<sub>3</sub>). The saponin (~35 mg) in EtOAc is chromatographed on Al<sub>2</sub>O<sub>3</sub> and eluted with \*C<sub>6</sub>H<sub>6</sub>/Et<sub>2</sub>O, and the residue on evaporation is recrystallised from Me<sub>2</sub>CO. [Mazur et al. *J Am Chem Soc* 82 5889 1960, Marker et al. *J Am Chem Soc* 69 2167 1947, *Beilstein* 19 III/IV 2581.]

**Hecogenin acetate** [915-35-5] **M 472.7, m 265-268 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{20} -6.0^{\circ}$  (c 1, CHCl<sub>3</sub>). Crystallise the acetate from MeOH. [*Beilstein* 19 IV 2583.]

**Hydrocortisone (Cortisol, 11 $\beta$ ,17 $\alpha$ ,21-trihydroxypregn-4-ene-3,20-dione)** [50-23-7] **M 362.5, m 212-213 $^{\circ}$ , 214-217 $^{\circ}$ , 218-221 $^{\circ}$ , 220-222 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{22} +167^{\circ}$  (c 1, EtOH). Recrystallise hydrocortisone from EtOH or isoPrOH. It is bitter tasting and has UV with  $\lambda_{\text{max}}$  at 242 nm (log  $\epsilon$  4.20). Its solubility at 25 $^{\circ}$  is: H<sub>2</sub>O (0.28%), EtOH (1.5%), MeOH (0.62%), Me<sub>2</sub>CO (0.93%), CHCl<sub>3</sub> (0.16%), propylene glycol (1.3%) and Et<sub>2</sub>O (0.35%). It gives an intense green colour with conc H<sub>2</sub>SO<sub>4</sub>. [Wendler et al. *J Am Chem Soc* 72 5793 1950, *Beilstein* 8 IV 3422.]

**Hydrocortisone acetate (21-acetoxy-11 $\beta$ ,17 $\alpha$ -trihydroxypregn-4-ene-3,20-dione)** [50-03-3] **M 404.5, m 218-221.5 $^{\circ}$ , 221-223 $^{\circ}$ , 222-225 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{25} +166^{\circ}$  (c 0.4, dioxane), **+150.7 $^{\circ}$  (c 0.5, Me<sub>2</sub>CO).** The acetate recrystallises from Me<sub>2</sub>CO/Et<sub>2</sub>O or aqueous Me<sub>2</sub>CO as *hygroscopic* monoclinic crystals. Its UV has  $\lambda_{\text{max}}$  at 242 nm ( $A_{1\text{cm}}^{1\%}$  390) in MeOH. Its solubility at 25 $^{\circ}$  is: H<sub>2</sub>O (0.001%), EtOH (0.45%), MeOH (0.04%), Me<sub>2</sub>CO (1.1%), CHCl<sub>3</sub> (0.5%), Et<sub>2</sub>O (0.15%), and it is very soluble in Me<sub>2</sub>NCHO. [Wendler et al. *J Am Chem Soc* 74 3630 1952; Antonucci et al. *J Org Chem* 18 7081 1953, *Beilstein* 8 IV 3424.]

**19-Hydroxy-4-androsten-3,17-dione** [510-64-5] **M 302.4, m 167-169 $^{\circ}$ , 168-170 $^{\circ}$ , 169-170 $^{\circ}$ , 172-173 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{20} +190^{\circ}$  (c 1, CHCl<sub>3</sub>). Recrystallise 19-hydroxy-4-androsten-3,17-dione from Me<sub>2</sub>CO/hexane or Et<sub>2</sub>O/hexane. It has UV with  $\lambda_{\text{max}}$  at 242nm in EtOH or MeOH. The *19-acetoxy* derivative has  $[\alpha]_{\text{D}}^{25} +185^{\circ}$  (CHCl<sub>3</sub>) and  $\lambda_{\text{max}}$  with 237.5nm in EtOH. [Ehrenstein & Dünneberger *J Org Chem* 21 774 1956, *Beilstein* 8 IV2162.]

**25-Hydroxycholesterol (cholest-5-en-3 $\beta$ ,25-diol)** [2140-46-7] **M 402.7, m 177-179 $^{\circ}$ , 178-180 $^{\circ}$ , 181.5-182.5 $^{\circ}$ ,**  $[\alpha]_{\text{D}}^{25} -39^{\circ}$  (c 1.05, CHCl<sub>3</sub>). 25-Hydroxycholesterol forms colourless needles from MeOH [Schwartz *Tetrahedron Lett* 22 4655 1981]. The *3 $\beta$ -acetoxy* derivative has **m 142-142.8 $^{\circ}$**  (from Me<sub>2</sub>CO) and  $[\alpha]_{\text{D}}^{25} -40.4^{\circ}$

(c 2, CHCl<sub>3</sub>). The 3 $\beta$ ,25-diacetoxy derivative has **m** 119-120.5° (from MeOH) and  $[\alpha]_D^{25}$  -35.5° (CHCl<sub>3</sub>). [Dauben & Bradlow *J Am Chem Soc* 72 4248 1950, Ryer et al. *J Am Chem Soc* 72 4247 1950, Beilstein 6 IV 6437.]

**18-Hydroxy-11-deoxycorticosterone (18,21-dihydroxypregn-4-en-3,20-dione tautomeric with 18,20-epoxy-20,21-dihydroxypregn-4-en-3-one)** [379-68-0] **M 346.5, m 168-170°, 171-173°, 191-195°, 200-205°,  $[\alpha]_D^{20}$  +151° (c 1, CHCl<sub>3</sub>).** Recrystallise 18-hydroxy-11-deoxycorticosterone from Et<sub>2</sub>O/Me<sub>2</sub>CO to give crystals **m** 200-205°. When it is recrystallised from Me<sub>2</sub>CO, it has **m** 191-195°. It has UV with  $\lambda_{\max}$  at 240nm. The 21-O-acetoxy-18-hydroxy derivative has **m** 158-159° (from Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub>), and the 21-O-acetoxy-18,20-epoxy derivative has **m** 149-154° (from Et<sub>2</sub>O). [Kahnt et al. *Helv Chim Acta* 38 1237 1955; Pappo *J Am Chem Soc* 81 1010 1959.]

**17 $\beta$ -Hydroxy-17 $\alpha$ -methyl-3-androsterone (Mestanolone)** [521-11-9] **M 304.5°, m 192-193°.** Dissolve mestanolone in Et<sub>2</sub>O, wash it with N NaOH, H<sub>2</sub>O, dry it (Na<sub>2</sub>SO<sub>4</sub>), evaporate and recrystallise it from EtOAc. The semicarbazone has **m** 235-236° (from EtOH). [Ruzicka et al. *Helv Chim Acta* 18 1487 1935.]

**17 $\alpha$ -Hydroxy-6 $\alpha$ -methylprogesterone (Medroxyprogesterone)** [520-85-4] **M 344.5, m 220°,  $[\alpha]_D^{25}$  +75° (CHCl<sub>3</sub>).** If it contains the *epi*-isomer (TLC), then dissolve it in CHCl<sub>3</sub>, bubble dry HCl gas to epimerise it, evaporate and recrystallise it from chloroform. The UV has  $\lambda_{\max}$  at 241nm ( $\epsilon$  16,150) in EtOH. The 17-acetate [71-58-9] crystallises from MeOH with **m** 207-208° and  $[\alpha]_D^{25}$  +61° (CHCl<sub>3</sub>). Its UV has  $\lambda_{\max}$  at 240nm ( $\epsilon$  15,950) in EtOH. [Babcock et al. *J Am Chem Soc* 80 2904 1958, Beilstein 8 IV 2212.]

**$\alpha$ -Hyodeoxycholic acid** [83-49-8] **M 392.6, m 196-197°,  $[\alpha]_{546}^{20}$  +8° (c 2, EtOH), pK<sub>Est</sub> ~4.9.** Crystallise  $\alpha$ -hyodeoxycholic acid from EtOAc or Me<sub>2</sub>CO. The K salt separates in needles from an alcoholic solution of the acid when an equivalent of KOMe is added (see lithocholic acid [434-13-9]). [Weiland & Gumlish *Hoppe Seyler's Z Physiol Chem* 215 18 1933, Windaus & Bohne *Justus Liebigs Ann Chem* 433 278 1923, Beilstein 10 III 1631.]

**Lanosterol** [79-63-0] **M 426.7, m 138-140°,  $[\alpha]_D^{20}$  +62.0° (c 1, CHCl<sub>3</sub>).** If very impure, then it should be acetylated, converted to the dibromide acetate [crystallised from EtOAc with slow cooling, **m** 168-170°,  $[\alpha]_D^{20}$  +214° (CHCl<sub>3</sub>)], de-brominated with Zn dust to give the acetate (below) which is recrystallised from 3-4 parts of Me<sub>2</sub>CO/MeOH (4:1) and hydrolysed as for stigmasterol (below). Recrystallise it from anhydrous MeOH. Dry it *in vacuo* over P<sub>2</sub>O<sub>5</sub> for 3 hours at 90°. The purity is checked by proton magnetic resonance. The acetate crystallises from MeOH with **m** 131-133°,  $[\alpha]_D^{25}$  +62° (c 1, CHCl<sub>3</sub>). [Block & Urech *Biochemical Preparations* 6 32 1958. van Tamelen et al. *J Am Chem Soc* 104 6479, 6480 1982, Beilstein 6 III 2880, 6 IV 4188.]

**Lithocholic acid (3 $\alpha$ -hydroxycholanic acid)** [434-13-9] **M 376.6, m 184-186°,  $[\alpha]_D^{25}$  +35° (c 1, EtOH), pK<sub>Est</sub> ~4.8.** Lithocholic acid can be purified by conversion to the rather insoluble Na or K salt by addition of the equivalent amount of aqueous NaOH or KOH, filtering off the alkali salt, washing it with ice cold H<sub>2</sub>O, dissolving it in the least volume of boiling H<sub>2</sub>O, acidifying with the dilute HCl (slight excess), filtering off the acid, washing with cold H<sub>2</sub>O and drying it thoroughly in a vacuum. Recrystallise it from Me<sub>2</sub>CO, EtOH or acetic acid. The methyl ester crystallises from MeOH, with 0.5 mol of MeOH, and has **m** 92-93°,  $[\alpha]_D^{25}$  +34° (MeOH). It has also been purified by recrystallisation from petroleum ether (b 40-60°) and, after chromatography on Al<sub>2</sub>O<sub>3</sub> in petroleum ether, gave a labile form **m** 92-93° which is transformed to the stable form **m** 125-126° after standing for 2 days in a vacuum desiccator. [Hoelm & Mason *J Am Chem Soc* 62 569 1940, Sarel & Yanuka *J Org Chem* 24 2018 1959, Beilstein 10 IV 785.]

**6- $\alpha$ -Methylprednisolone (Medrol, 11 $\beta$ ,17-21-trihydroxy-6 $\alpha$ -methylpregna-1,4-dien-3,20-dione)** [83-43-2] **M 347.5, m 226-237°, 228-237°, 240-242°,  $[\alpha]_D^{24}$  +91° (c 0.5, dioxane).** Recrystallise medrol from EtOAc. The UV has  $\lambda_{\max}$  at in 95% EtOH 243nm ( $\epsilon$  14,875). The 21-acetoxy derivative has **m** 205-208° (from EtOAc), and  $[\alpha]_D^{24}$  +95° (c 1, CHCl<sub>3</sub>). [Spero et al. *J Am Chem Soc* 78 6213 1956; Fried et al. *J Am Chem Soc* 81 1235 1959; <sup>1</sup>H NMR: Slomp & McGarvey *J Am Chem Soc* 81 2200 1959, Beilstein 8 IV 3498.]

**17 $\alpha$ -Methyltestosterone (Android, Mesterone)** [58-18-4] M 302.5, m 162-168 $^{\circ}$ , 164-165 $^{\circ}$ ,  $[\alpha]_{D}^{20} +87^{\circ}$  (c 1, dioxane),  $[\alpha]_{D}^{20} +82.3^{\circ}$  (c 1, EtOH)  $A_{1cm}^{1\%}$  241nm is 495-530 (EtOH). This anabolic steroid is crystallised from hexane or hexane/\*benzene. It has  $E_{1cm}^{1\%}$  495-530 at 241nm (EtOH). The colour reaction with 2,4-dinitrophenyl hydrazine is used for assaying it. [Gornall & Macdonald *J Biol Chem* **201** 279 1953.] In another colour reaction the sterone (1mg) in acetic acid (0.2ml) + 88% H<sub>3</sub>PO<sub>4</sub> (2ml) is allowed to stand for 1 hour when it becomes fluorescent. After 1 hour it is diluted with acetic acid (~3ml) and provides a strong yellow fluorescence with the intensity of 50-100 times that of estrone. [Stuart & Stuckey *J Pharm Pharmacol* **1** 130 1949, Openauer *Rec Trav Chim Pays Bas* **56** 137 137, *Beilstein* **8** IV 1010.]

**Norcholanic acid (5 $\beta$ -24-norcholan-23-oic acid)** [511-18-2] M 346.5, m 175.5-176.5 $^{\circ}$ , 186 $^{\circ}$ ,  $[\alpha]_{D}^{20} +32^{\circ}$  (EtOH), pK<sub>Est</sub> ~4.8. Recrystallise the acid from AcOH. The *methyl ester* has m 74 $^{\circ}$  (needles from MeOH), and the *ethyl ester* has m 65-66 $^{\circ}$ . [Yanuka et al. *Tetrahedron Lett* 1725 1968, *Beilstein* **9** III 2652, **10** IV 2083.]

**$\alpha$ -Oestradiol** See estradiol above [57-91-0].

**$\beta$ -Oestradiol-3-benzoate** See estradiol 3-benzoate above [50-50-0].

**Ouabain [3-{(6-deoxy- $\alpha$ -L-mannopyranosyl)oxy}-1,5,11a,14,19-pentahydroxycard-20(22)-enolide. G-Strophanthin, Acocantherine]** [630-60-4, 312619-45-7 (hydrate), 11018-89-6 (octahydrate)] M 728.8 (8 H<sub>2</sub>O), m 190 $^{\circ}$ (dec), 200-202 $^{\circ}$ (dec),  $[\alpha]_{546}^{20} -30^{\circ}$  (c 1, H<sub>2</sub>O). It crystallises from water as the *octahydrate*. Dry it at 130 $^{\circ}$ . It decomposes at 190 $^{\circ}$  when dry. Store it in the dark as it is light sensitive, but it is stable in air. Its solubility (g/100ml) in H<sub>2</sub>O is 1.3 (~25 $^{\circ}$ ), 20 (~100 $^{\circ}$ ), and in EtOH it is 1.0 (~25 $^{\circ}$ ) and 12.5 (~78 $^{\circ}$ ). It is highly **TOXIC** as it is an inhibitor of cation transport and of Na<sup>+</sup> and K<sup>+</sup> ATPase. [*Beilstein* **18/5** V 625.]

**Pancuronium bromide (2 $\beta$ ,16 $\beta$ -dipiperidino-5 $\alpha$ -androstan-3 $\alpha$ ,17 $\beta$ -diol diacetate dimetho-bromide)** [15500-66-0] M 732.7, m 212-215 $^{\circ}$ , 215 $^{\circ}$ . The bromide forms odourless crystals with a bitter taste which are purified through acid-washed Al<sub>2</sub>O<sub>3</sub> and eluted with isoPrOH/EtOAc (3:1) to remove impurities (e.g. the monomethobromide) and eluted with isoPrOH to give the pure dibromide which is recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/Me<sub>2</sub>CO or isoPrOH/Me<sub>2</sub>CO. It is soluble in H<sub>2</sub>O (10%) and CHCl<sub>3</sub> (3.3%) at 20 $^{\circ}$ . It is a non-depolarising muscle relaxant. [Buckett et al. *J Med Chem* **16** 1116 1973.]

**Prednisolone acetate (21-acetoxypregna-1,4-diene-11 $\beta$ -17 $\alpha$ -diol-3,20-dione)** [52-21-1] M 402.5, m 237-239 $^{\circ}$ , 240-242 $^{\circ}$ , 240-243 $^{\circ}$ , 244 $^{\circ}$ ,  $[\alpha]_{D}^{20} +116^{\circ}$  (c 1, dioxane). Recrystallise prednisolone acetate from EtOH, Me<sub>2</sub>CO, Me<sub>2</sub>CO/hexane, and it has UV with  $\lambda_{max}$  at 243nm in EtOH. [Joly et al. *Bull Soc Chim Fr* 366 1958; Herzog et al. *J Am Chem Soc* **77** 4781 1955, *Beilstein* **8** IV 3468.]

**Prednisone (17,21-dihydroxypregna-1,4-diene-3,11,20-trione, deltacortisone)** [53-03-2] M 358.5, m 238 $^{\circ}$  (dec),  $[\alpha]_{D}^{20} +172^{\circ}$  (c 0.5, dioxane),  $\lambda_{max}$  238nm (log  $\epsilon$  4.18) in MeOH. Crystallise prednisone from acetone/hexane, then recrystallise it from Me<sub>2</sub>CO. The *monoacetate* crystallises from Me<sub>2</sub>CO/hexane with m 227-233 $^{\circ}$ (dec),  $[\alpha]_{D}^{25} +186^{\circ}$  (c 1, dioxane), and the *diacetate* crystallises from Me<sub>2</sub>CO/hexane with m 219-221 $^{\circ}$ (dec),  $[\alpha]_{D}^{25} +125^{\circ}$  (c 1, CHCl<sub>3</sub>). [Hertzog et al. *Tetrahedron* **18** 581 1962, *Beilstein* **8** IV 3531.]

**5 $\alpha$ -Pregnane (allopregnane)** [641-85-0] M 288.5, m 84.5-85 $^{\circ}$ ,  $[\alpha]_{D}^{20} +21.7^{\circ}$  (c 1.3, CHCl<sub>3</sub>). Recrystallise 5 $\alpha$ -pregnane several times from Me<sub>2</sub>CO. The melting point is lowered (e.g. to 50-71 $^{\circ}$ ) if it is contaminated with the 5 $\beta$ -isomer (see below). [Butenandt et al. *Chem Ber* **64** 2529 1931, Steiger & Reichstein *Helv Chim Acta* **21** 161 1938, *Beilstein* **5** III 1120, 1121, 1125.]

**5 $\beta$ -Pregnane** [481-26-5] M 288.5, m 82-83 $^{\circ}$ , 83.5 $^{\circ}$ ,  $[\alpha]_{D}^{20} +21.2^{\circ}$  (c 0.75, CHCl<sub>3</sub>). Crystallise 5 $\beta$ -pregnane from MeOH or Me<sub>2</sub>CO. The mixed melting point with allopregnane (above) is ~50-71 $^{\circ}$ . [Butenandt et al. *Chem*

*Ber* 64 2529 1931, Steiger & Reichstein *Helv Chim Acta* 21 161 1938.]

**5 $\alpha$ -Pregnane-3 $\alpha$ ,20 $\beta$ R-diol** [566-57-5] M 320.5, m 207-209°, [ $\alpha$ ]<sub>D</sub><sup>26</sup> +12° (c 1.1, CHCl<sub>3</sub>). Crystallise 5 $\alpha$ -pregnane-3 $\alpha$ ,20 $\alpha$ -diol from Me<sub>2</sub>CO. Its *diacetate* [6170-22-5] M 404.5, has two melting points 134-140° and 152.4°, and [ $\alpha$ ]<sub>D</sub><sup>26</sup> +40.5° (c 0.8, CHCl<sub>3</sub>). It is a progesterone metabolite in urine during pregnancy. [See references below.]

**5 $\alpha$ -Pregnane-3 $\alpha$ ,20 $\alpha$ S-diol** [566-58-5] M 320.5, m 248°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +17° (c 0.15, EtOH). Crystallise 5 $\alpha$ -pregnane-3 $\alpha$ ,20 $\alpha$ -diol from EtOH. Its *diacetate* [6003-18-5] has m 142° and [ $\alpha$ ]<sub>D</sub><sup>20</sup> +18° (c 0.4, \*C<sub>6</sub>H<sub>6</sub>). It is a progesterone metabolite in urine during pregnancy. [See references below.]

**5 $\beta$ -Pregnane-3 $\alpha$ ,20 $\beta$ R-diol** [80-91-1] M 320.5, m 239°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +10° (c 1, EtOH). Crystallise pregnane-diol (which is abundant in the urine of pregnant women) from EtOH or Me<sub>2</sub>CO and dry it *in vacuo*. It can be oxidised to progesterone (see below) and it is not precipitated by digitonin. Its *diacetate* [6100-28-3] has m 112-113° and [ $\alpha$ ]<sub>D</sub><sup>26</sup> +60° (c 1, CHCl<sub>3</sub>). It is a progesterone metabolite in urine during pregnancy. [Marian *Biochem J* 23 1090 1929, Johnson et al. *J Chem Soc* 1302 1954, Mattox et al. *J Org Chem* 32 708 1967, *Beilstein* 6 III 4778, 6 IV 6111.]

**5 $\beta$ -Pregnane-3 $\alpha$ ,20 $\alpha$ S-diol** [80-92-2] M 320.5, m 243-244°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +31° (c 1, EtOH). Crystallise the diol from acetone. The *diacetate* [1174-69-2] crystallises from petroleum ether with m 180° (also 182-183°) and [ $\alpha$ ]<sub>D</sub><sup>20</sup> +35° (c 1.1, CHCl<sub>3</sub>). [Marian *Biochem J* 23 1090 1929, Fish et al. *J Biol Chem* 143 716 1942, Hieschmann *J Biol Chem* 140 797 1941, Johnson et al. *J Chem Soc* 1302 1954, Glick et al. *J Org Chem* 27 3121 1962, *Beilstein* 6 III 4778, 6 IV 6111.]

**Progesterone** [57-83-0] M 314.5, m 128.5°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +230° (c 1, EtOH), +214.7° (c 1.3, Me<sub>2</sub>CO). The  $\alpha$ -form crystallises from EtOH with m 127-131°. The  $\beta$ -form crystallises from petroleum ether or aqueous petroleum ether/aqueous Et<sub>2</sub>O with m 119-120° or 121°. It also crystallises from Et<sub>2</sub>O, Me<sub>2</sub>CO/EtOAc, MeOH, aqueous Et<sub>2</sub>O, aqueous MeOH, wet petroleum ether, Et<sub>2</sub>O/petroleum ether, petroleum ether/\*C<sub>6</sub>H<sub>6</sub>, Et<sub>2</sub>O/pentane and isopropyl ether. The UV has  $\lambda_{\max}$  is at 240nm with log  $\epsilon$  4.25 (EtOH). [Wintersteiner & Allen *J Biol Chem* 107 321 1934, *Beilstein* 7 III 3648, 7 IV 2395.]

**Rubijervine (solanid-5-ene-3 $\beta$ -12 $\alpha$ -diol)** [79-58-3] M 413.6, m 242-244°, 240-246°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> +19° (EtOH), [ $\alpha$ ]<sub>D</sub><sup>27.5</sup> +20° (c 1.04, MeOH), pK<sub>Est</sub> ~7.0. Rubijervine crystallises from 95% EtOH as colourless rods. It has solvent of crystallisation and is dried at 120°/2mm. It is precipitated by digitonin. The *hydrobromide* crystallises from MeOH/Me<sub>2</sub>CO with m 265-270°(dec). The *diacetate* crystallises from MeOH with m 160-163°. The *3-benzoate* gives colourless prisms from \*C<sub>6</sub>H<sub>6</sub> with m 156-159° and [ $\alpha$ ]<sub>D</sub><sup>27.5</sup> +22° (c 1.6, CHCl<sub>3</sub>). [Pelletier & Locke *J Am Chem Soc* 79 4531 1957, Jacobs & Craig *J Biol Chem* 148 41 1943, *Beilstein* 21 III/IV 2310.]

**$\beta$ -Sitosterol (stigmast-5-ene-3 $\beta$ -ol)** [83-46-5] M 414.7, m 136.5-137.5°, 140°, [ $\alpha$ ]<sub>D</sub><sup>20</sup> -42° (c 2, CHCl<sub>3</sub>). Crystallise  $\beta$ -sitosterol from EtOH, MeOH, Me<sub>2</sub>CO, Me<sub>2</sub>CO/EtOH or Me<sub>2</sub>CO/MeOH. It has also been purified by zone melting. The *acetate* crystallises from MeOH or EtOH as plates with m 127-128° and [ $\alpha$ ]<sub>D</sub><sup>20</sup> -41° (c 2, CHCl<sub>3</sub>). The *benzoate* crystallises from EtOH as needles with m 146-147° and [ $\alpha$ ]<sub>D</sub><sup>20</sup> -13.8° (c 2, CHCl<sub>3</sub>). [Fujimoto & Jacobson *J Org Chem* 29 3377, 3381 1964, Shoppee *J Chem Soc* 1043 1948, Heilbron et al. *J Chem Soc* 344, 347 1941, *Beilstein* 6 III 2696.]

**Smilagenin (25R-spirostan-3 $\beta$ -ol, isosarsapogenin)** [126-18-1] M 416.6, m 185°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -69°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -80° (c 0.3, CHCl<sub>3</sub>). Chromatograph smilagenin on active Al<sub>2</sub>O<sub>3</sub> and elute with \*C<sub>6</sub>H<sub>6</sub>, then recrystallise it from Me<sub>2</sub>CO, aqueous EtOH (m 187-188°) or MeOH. The *acetate* crystallises from MeOH with m 152° and [ $\alpha$ ]<sub>D</sub><sup>25</sup> -59.6°, [ $\alpha$ ]<sub>D</sub><sup>25</sup> -68.9° (c 0.25, CHCl<sub>3</sub>). [Askew *J Chem Soc* 1399 1936 and 1402 1936, Scheer et al. *J Am Chem Soc* 77 641 1955, *Beilstein* 19 III/IV 826.]

**Solanidine (solanid-5-en-3 $\beta$ -ol)** [80-78-4] M 397.6, m 218-219°(sublimes), [ $\alpha$ ]<sub>D</sub><sup>20</sup> -29° (c 0.5, CHCl<sub>3</sub>), pK<sup>15</sup>

**6.66.** Solanidine crystallises from  $\text{CHCl}_3/\text{MeOH}$ , aqueous EtOH or aqueous MeOH as needles. TLC on  $\text{Al}_2\text{O}_3$  plates using  $\text{CH}_2\text{Cl}_2/\text{MeOH}$  (98:2) gives a spot at  $R_F$  0.47. The *hydrochloride* crystallises from aqueous EtOH with **m** 345°(dec). The *acetate* crystallises from EtOH with **m** 208°. [Schreiber & Roensch *Tetrahedron* **21** 6451/965, Kessar et al. *Tetrahedron* **27** 2153 1971, Reichstein & Reich *Ann Rev Biochem* **15** 155 1946, Beilstein **21** III/IV 1398, **27** III/IV 2000.]

**$\alpha$ -Solanine (solan-5-en-3 $\beta$ -yl-[*O*<sup>3</sup>- $\beta$ -D-glucopyranosyl-*O*<sup>2</sup>- $\alpha$ -L-rhamnopyranosyl- $\beta$ -D-galactopyrano-side])**  
See “Carbohydrates” in this chapter.

**Solasodine (22 $\alpha$ ,25*R*-spirosol-5-en-3 $\beta$ -ol) [126-17-0] M 413.6, m 202°,  $[\alpha]_D^{25}$  -100° (c 2, MeOH), **pK**<sup>25</sup> 7.7.** Solasodine crystallises (as *monohydrate*) from MeOH as lustrous plates (on heating, the plates change to needles as they melt and resolidify in needles), or aqueous 80% EtOH, and sublimes at high vacuum. After recrystallisation from  $\text{H}_2\text{O}$ , **m** 198-199°, then from  $\text{Me}_2\text{CO}/\text{H}_2\text{O}$ , **m** 199-201°, it has  $[\alpha]_D^{25}$  -109.3° (c 0.581,  $\text{CHCl}_3$ ). On TLC with silica gel G ( $^*\text{C}_6\text{H}_6/\text{absolute EtOH}$ , 8:2) it has  $R_F$  0.45. The IR (KBr) has  $\nu_{\text{max}}$  at 10.3, 10.4, 11.2, 11.5 $\mu$  (azaooxaspirane bands). [Schreiber & Rönsh *Tetrahedron* **20** 1939 1964, Uhle *J Org Chem* **27** 656 1962, Kessar et al. *Tetrahedron* **27** 2153 1971, Beilstein **27** III/IV 2000.]

**Stigmasterol (3 $\beta$ -hydroxy-24-ethylcholesta-5,22-diene) [83-48-7] M 412.7, m 170°,  $[\alpha]_D^{25}$  -51° ( $\text{CHCl}_3$ ),  $[\alpha]_{546}^{20}$  -59° (c 2,  $\text{CHCl}_3$ ).** Stigmasterol is best purified *via* the *tetrabromide-acetate*. The impure sterol (3g) is acetylated with  $\text{Ac}_2\text{O}$  (60ml) by refluxing for 1.5 hour. The mixture is cooled at 20° for 1 hour, and the crude acetate is collected. The acetate (3g) in  $\text{Et}_2\text{O}$  (30ml) is then treated with  $\text{Br}_2/\text{AcOH}$  (38ml, from 5g  $\text{Br}_2$  in 100ml AcOH), and after cooling at 6° overnight, the tetrabromoacetate is filtered off and washed with  $\text{Et}_2\text{O}$ . After six recrystallisations from  $\text{CHCl}_3/\text{MeOH}$  the *tetrabromoacetate* has **m** 194-196°. This product (1g) in AcOH (12ml) and Zn dust (1g) is refluxed for 1.5 hours, filtered hot, diluted with  $\text{H}_2\text{O}$  (30m L) and extracted with  $\text{Et}_2\text{O}$ . The extract is washed with dilute aqueous sodium sulfite, then  $\text{H}_2\text{O}$ , the extract is dried ( $\text{Na}_2\text{SO}_4$ ) and the stigmasterol acetate (~550mg) is recrystallised (4x) from EtOH and twice from MeOH/ $\text{CHCl}_3$  (2:1) to give the *acetate* with **m** 139-148°. This acetate (400mg) is hydrolysed in boiling 10% alcoholic KOH (1ml) for 1 hour. Then  $\text{H}_2\text{O}$  (30ml) is added and the mixture is extracted with  $\text{Et}_2\text{O}$ . The extract is washed with aqueous  $\text{Na}_2\text{CO}_3$ , then  $\text{H}_2\text{O}$ , the solvent is distilled off and the residue is recrystallised (3x) from 95% EtOH to give ~110mg of pure stigmasterol. It is dried in a vacuum over  $\text{P}_2\text{O}_5$  for 3 hours at 90°. The purity is checked by NMR. The *acetate* crystallises from MeOH with **m** 145°,  $[\alpha]_D^{25}$  -56° (c 2,  $\text{CHCl}_3$ ). [Byerrum & Ball *Biochemical Preparations* **7** 86 1959, Thornton et al. *J Am Chem Soc* **62** 2006 1940, Colin et al. *Anal Chem* **51** 1661 1979, Beilstein **6** IV 4170.]

**Taurocholic acid (3 $\alpha$ , 7 $\alpha$ ,12 $\alpha$ -trihydroxy-5 $\beta$ -cholestan-24-oic acid *N*-(2-sulfoethyl) amide Na salt, *N*-coloyotaurine) [81-24-3] M 515.6, m ~125°(dec),  $[\alpha]_D$  +38.8° (c 2, EtOH), **pK**<sup>25</sup> 1.4.** The acid is present in bile and is isolated as an amorphous pale yellow powder. Crystallise it from EtOH/EtOAc/ $\text{Et}_2\text{O}$  (amorphous **m** 125°) or EtOH/ $\text{Et}_2\text{O}$  [Josephson *Biochem J* **29** 1484 1935]. The anhydrous acid is hygroscopic, freely soluble in  $\text{H}_2\text{O}$  and EtOH but insoluble in  $\text{H}_2\text{O}$  and EtOAc. It is hydrolysed by acids or alkalis to cholic acid and taurine. The *sodium salt (hydrate)* [312693-83-7, 345909-26-4 ( $x\text{H}_2\text{O}$ )] crystallises from aqueous EtOH/ $\text{Et}_2\text{O}$  with **m** 231-232°,  $[\alpha]_D^{23}$  +23.6° (c 2.5,  $\text{H}_2\text{O}$ ); and has the UV  $\lambda_{\text{max}}$  ( $\text{H}_2\text{SO}_4$ ) at 303, 389 and 480nm. [cf Cortese *J Am Chem Soc* **59** 2532 1937, Beilstein **10** III 2177, **10** IV 2078.]

**Taurodeoxycholic acid sodium salt monohydrate (3 $\alpha$ , 12 $\alpha$ -dihydroxy-5 $\beta$ -cholestan-24-oic acid *N*-(2-sulfoethyl) amide Na salt monohydrate, *N*-[desoxycholy]taurine Na salt  $\text{H}_2\text{O}$ ) [1180-95-6, 207737-97-1 (*x H}\_2\text{O}*)] M 539.7, m 171-175°,  $[\alpha]_D^{23}$  +37° (c 1,  $\text{H}_2\text{O}$ ), **pK**<sup>25</sup> 1.4 (free acid).** The Na salt is dissolved in the smallest volume of  $\text{H}_2\text{O}$ , a saturated solution of aqueous NaCl/ $\text{Et}_2\text{O}$  is added and the mixture is stored at 0° for 24 hours. Then shake the mixture well, keep it in the cold for another day and filter the crystals with gentle shaking, wash them with ice-cold saturated aqueous NaCl saturated with  $\text{Et}_2\text{O}$ , dry them over  $\text{CaCl}_2$  and extract them with absolute EtOH. Add ~10ml of  $\text{H}_2\text{O}$  to the solid followed by enough  $\text{Et}_2\text{O}$  to incipient cloudiness. Store it overnight at 0°. Add ice-cold  $\text{Et}_2\text{O}$  to make 250ml, collect the crystals, wash them with  $\text{Et}_2\text{O}$  then petroleum ether and dry them in air. The purification can be repeated with NaCl and  $\text{Et}_2\text{O}$  with ~85% recovery. NB: precipitation will not occur unless enough  $\text{H}_2\text{O}$  is present. Its solubility in  $\text{H}_2\text{O}$  is 10%. [Cortese *J Am*

*Chem Soc* **59** 2532 1937.] The *free acid* has **m** 141-144°. [Norman *Ark Kemi* **8** 331 1956.] It forms mixed micelles and solubilises some membrane proteins [Hajjar et al. *J Biol Chem* **258** 192 1983]. [*Beilstein* **10** IV 1611.]

**Testosterone (17- $\beta$ -hydroxyandrost-4-ene-3-one)** [58-22-0] **M 288.4, m 155°,  $[\alpha]_{546}^{20} +130^\circ$  (c 1, dioxane).** Crystallise testosterone from aqueous acetone, hexane or isoPrOH. The long needles that separated from EtOH/AcOH were used for X-ray crystallography [Roberts et al. *J Chem Soc Perkin Trans II* 1978 1973.] The *acetate* [1045-69-8] crystallises from MeOH or aqueous Me<sub>2</sub>CO, with **m** 140-141° and  $[\alpha]_{\text{D}}^{20} +87.8^\circ$  (c 1, EtOH). [Ruzicka et al. *Helv Chim Acta* **18** 1478 1935 and **19** 99, 842 1936, *Beilstein* **8** IV 974.]

**Testosterone propionate** [57-85-2] **M 344.5, m 118-122°,  $[\alpha]_{546}^{20} +102^\circ$  (c 2, dioxane).** Crystallise the propionate from aqueous EtOH, or Et<sub>2</sub>O/petroleum ether (**m** 121°), and its UV has  $\lambda_{\text{max}}$  at 240nm (EtOH), and  $[\alpha]_{546}^{20} +114^\circ$  (c 1, CHCl<sub>3</sub>). Also purify it by HPLC. [Ercol & de Ruggieri *J Am Chem Soc* **75** 650, 652 1953, polymorphism: Brandstätter-Kuhnert & Kofler *Microchim Acta* 847, 850 1959, *Beilstein* **8** IV 977.]

**Ursodiol (ursodeoxycholic acid, 3 $\alpha$ ,7 $\beta$ -dihydroxy-5 $\beta$ -cholan-24-oic acid, 7 $\beta$ -hydroxylithocholic acid, ursodeoxycholic acid)** [128-13-2] **M 392.5, m 203°,  $[\alpha]_{\text{D}}^{20} +60^\circ$  (c 0.2, EtOH),  $\text{pK}_{\text{Est}} \sim 4.8$ .** Recrystallise ursodiol from wet Et<sub>2</sub>O, EtOH or EtOH/MeOH. It is almost insoluble in H<sub>2</sub>O, sparingly soluble in Et<sub>2</sub>O, very slightly soluble in CHCl<sub>3</sub> but freely soluble in AcOH. The *diformate* has **m** 170°, and the *diacetate* has **m** 100-102°. It is an anticholelithogenic drug. [Iwasaki *Hoppe Seyler's Z Physiol Chem* **244** 181, 183 1936, Ward et al. *Drugs* **27** 95 1984, *Beilstein* **10** III 1635, **10** IV 1604.]

## MISCELLANEOUS COMPOUNDS (including biologically useful reagents, low-molecular-weight bioactive substances, antibiotics, coenzymes, vitamins, lipids, phospholipids, nucleosides, nucleotides and polynucleotides)

**Acetoacetyl coenzyme A trisodium salt trihydrate** [102029-52-7] **M 955.6, pK<sub>1</sub> 4.0 (NH<sub>2</sub>), pK<sub>2</sub> 6.4 (PO<sub>4</sub><sup>-</sup>)**. Purification can be carried out by passage through a DEAE-cellulose formate column, then through a Dowex 50 (H<sup>+</sup>) column to remove Na ions, concentrated by lyophilisation and redissolved in H<sub>2</sub>O. It is commercially available as a solution of 0.05g/ml of H<sub>2</sub>O. The concentration of acetoacetylcoenzyme A is determined by the method of Stern et al. [*J Biol Chem* **221** 15 1956]. It is stable at pH 7-7.5 for several hours at 0° (half-life *ca* 1-2 hours). At room temperature it is hydrolysed in *ca* 1-2 hours at pH 7-7.5. At pH 1.0/20° it is more stable than at neutrality. A solution of the trisodium salt (0.05g/ml H<sub>2</sub>O) adjusted to pH 5 with 2N NaOH can be stored frozen for several weeks. It is stable at pH 2-3/-17° for at least 6 months. [Hersch & Jencks *J Biol Chem* **242** 3468 1967, Clikenbeard et al. *J Biol Chem* **250** 3108 1975, Simon & Shemin *J Am Chem Soc* **75** 2520 1953, Moffatt & Khorana **81** 1265 1959, Salem et al. *Biochem J* **258** 563 1989, Beilstein **26** III/IV 3668.]

**Acetylcholine bromide** [66-23-9] **M 226.1, m 143°, 146°**. The bromide is a *hygroscopic* solid, but less so than the hydrochloride salt. It crystallises from EtOH as prisms. Some hydrolysis occurs in boiling EtOH, particularly if it contains some H<sub>2</sub>O. It can also be recrystallised from EtOH or MeOH by adding dry Et<sub>2</sub>O. [Heilbronn *Acta Chem Scand* **12** 1492 1958, Beilstein **4** IV 1446.]

**Acetylcholine chloride** [60-31-1] **M 181.7, m 148-150°, 151°**. It is very soluble in H<sub>2</sub>O (>10%), and is very *hygroscopic*. If pasty, dry it in a vacuum desiccator over H<sub>2</sub>SO<sub>4</sub> until a solid residue is obtained. Dissolve this in absolute EtOH, filter it and add dry Et<sub>2</sub>O, when the hydrochloride separates. Collect by filtration and store it under very dry conditions. [Jones & Major *J Am Chem Soc* **52** 307 1930.] The *chloroplatinate* crystallises from hot H<sub>2</sub>O in yellow needles and can be recrystallised from 50% EtOH, **m 242-244°** [Dudley *Biochem J* **23** 1069 1929]; other **m** given is 256-257°. The *perchlorate* crystallises from EtOH as prisms **m 116-117°**. [*J Am Pharm Assocn* **36** 272 1947, Beilstein **4** IV 1446.]

**N<sup>4</sup>-Acetylcytosine** [14631-20-0] **M 153.1, m >300°, 326-328°, pK<sub>Est(1)</sub> ~1.7, pK<sub>Est(2)</sub> ~10.0**. If TLC or paper chromatography shows that it contains unacetylated cytosine, then reflux it in Ac<sub>2</sub>O for 4 hours, cool at 3-4° for a few days, collect the crystals, wash them with cold H<sub>2</sub>O, then EtOH and dry at 100°. It is insoluble in EtOH and dissolves in H<sub>2</sub>O with difficulty, but crystallises in prisms from hot H<sub>2</sub>O. It is hydrolysed by 80% aqueous AcOH at 100°/1 hour. [UV: Brown et al. *J Chem Soc* 2384 1956, Codington et al. *J Am Chem Soc* **80** 5164 1958.] It forms an Hg salt [Fox et al. *J Am Chem Soc* **79** 5060 1957]. [Beilstein **25** III/IV 3657.]

**N-Acetylhistamine [N-(2-1{3}H-imidazol-4-yl)ethylacetamide]** [673-49-4] **M 153.2, m 147-148°, 148-149°, pK<sup>25</sup> 6.99**. It is purified by recrystallisation from Et<sub>2</sub>O/EtOH or EtOAc/EtOH (needles), and dried *in vacuo*. It sublimes at 148°/0.05mm. It is slightly soluble in H<sub>2</sub>O and EtOH, soluble in Me<sub>2</sub>CO and very soluble in Et<sub>2</sub>O. The *nitrate* has **m 170°**, the *picrate* is dimorphic with **m 169-171°** (from EtOH/Et<sub>2</sub>O) and **m 181-183°** (from EtOH). [Tabor & Mosettig *J Biol Chem* **80** 703 1949, Nagarajan et al. *Indian J Chem, Sect B* **16** 629, 633 1977, Beilstein **25** II 304, **25** III/IV 2053, **25/9** V 523.]

**O-Acetyl-β-methylcholine chloride [Methacholine chloride, Amechol, Provocholine, 2-acetoxypropylammonium chloride]** [62-51-1] **M 195.7, m 170-173°, 172-173°**. It forms white *hygroscopic* needles from Et<sub>2</sub>O and is soluble in H<sub>2</sub>O, EtOH and CHCl<sub>3</sub>. It decomposes readily in alkaline solutions and slowly in H<sub>2</sub>O. It should be handled and stored in a dry atmosphere. The *bromide* is less *hygroscopic*, and the *picrate* has **m 129.5-131°** (from EtOH). [racemate: Annis & Ely *Biochem J* **53** 34 1953, IR of iodide: Hansen *Acta Chem Scand* **13** 155 1959, Beilstein **4** IV 1670.]

**N-Acetyl neuraminic acid (NANA, O-Sialic acid, 5-acetamido-3,5-dideoxy-D-glycero-D-galacto-2-nonulosonic acid, lactaminic acid)** [131-48-6] **M 309.3, m 159°(dec), 181-183°(dec), 185-187°(dec), [α]<sub>D</sub><sup>25</sup> -33° (c 2, H<sub>2</sub>O), pK<sup>25</sup> 2.6**. A Dowex-1x8 (200-400 mesh) in the formate form is used, and is prepared by

washing with 0.1M NaOH, then 2N sodium formate; excess formate is removed by washing with H<sub>2</sub>O. *N*-Acetylneuraminic acid in H<sub>2</sub>O is applied to this column, washed with H<sub>2</sub>O, then eluted with 2N formic acid at a flow rate of 1ml/minute. Fractions (20ml) are collected and tested (Bial's orcinol reagent, *cf Biochemical Preparations* 7 1 1960). NANA elutes at a formic acid molarity of 0.38, and the Bial positive fractions are collected and lyophilised. The residue is recrystallised from aqueous AcOH: suspend 1.35g of residue in AcOH, heat rapidly to boiling, add H<sub>2</sub>O dropwise until the suspension dissolves (do not add excess H<sub>2</sub>O), filter hot and then keep at +5° for several hours until crystallisation is complete. Collect NANA and dry it in a vacuum over P<sub>2</sub>O<sub>5</sub>. *Alternatively*, dissolve 1.35g of NANA in 14ml of H<sub>2</sub>O, filter, add 160ml of MeOH followed by 360ml of Et<sub>2</sub>O. Then add petroleum ether (b 40-60°) until heavy turbidity. Cool at 20° overnight. The yield of NANA is *ca* 1.3g. Dry it over P<sub>2</sub>O<sub>5</sub> at 100°/1mm to constant weight. It mutarotates in Me<sub>2</sub>SO: [α]<sub>D</sub><sup>20</sup> -115° (after 7 minutes) to -32° (after 24 hours). It is commercially available as an aqueous solution (0.01g/ml). [IR and synthesis: Cornforth et al. *Biochem J* 68 57 1958; Zillikin & O'Brien *Biochemical Preparations* 7 1 1960; <sup>13</sup>C NMR and 1-<sup>13</sup>C synthesis: Nguyen & Perry *J Org Chem* 43 551 1978; Danishefski & DeNinno *J Org Chem* 51 2615 1986; Gottschalk, *The Chemistry and Biology of Sialic Acids and Related Substances*, Cambridge University Press, London, 1960, *Beilstein* 4 IV 3288.]

***N*-Acetyl penicillamine (*N*-acetyl-3-mercapto-D-valine)** [D -15537-71-0, DL -59-53-0] **M 191.3, m 183°, 186-187° (DL-form), 189-190° (D-form), D-form** [α]<sub>D</sub><sup>25</sup>+18° (c 1, 50% EtOH), pK<sub>Est(1)</sub> ~3.0 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~8.0 (SH). Both forms are recrystallised from hot H<sub>2</sub>O. A pure sample of the D-form is obtained after five recrystallisations. [Crooks in *The Chemistry of Penicillin* Clarke, Johnson and Robinson eds, Princeton University Press, 470 1949, Review: Chain et al. *Antibiotics* (Oxford University Press) 2 1949, *Beilstein* 4 III 1662.]

***p*-Acetylphenyl sulfate potassium salt**, [38533-41-4] **M 254.3, m dec on heating, pK<sub>Est</sub> ~2.1**. Purify the salt by dissolving it in the minimum volume of hot water (60°) and adding EtOH, with stirring, then leave at 0° for 1 hour. The crystals are filtered off and recrystallised from H<sub>2</sub>O until free of Cl<sup>-</sup> and SO<sub>4</sub><sup>2-</sup> ions. Dry it in a vacuum over P<sub>2</sub>O<sub>5</sub> at room temperature. It is a specific substrate for arylsulfatases which hydrolyse it to *p*-acetylphenol [UV has λ<sub>max</sub> at 327nm (ε 21700 M<sup>-1</sup>cm<sup>-1</sup>)] [Milsom et al. *Biochem J* 128 331 1972].

**S-Acetylthiocholine bromide** [25025-59-6] **M 242.2, m 217-223°(dec)**. It is a *hygroscopic* solid which can be recrystallised from ligroin/EtOH (1:1), dried and kept in a vacuum desiccator. Crystallisation from \*C<sub>6</sub>H<sub>6</sub>/EtOH gives **m** 227° or from propan-1-ol the **m** is 213°. [Hansen *Acta Chem Scand* 11 537 1957, Heilbronn *Acta Chem Scand* 12 1481 1958, *Beilstein* 4 IV 1585.]

**S-Acetylthiocholine chloride** [6050-81-3] **M 197.7, m 172-173°**. The chloride can be purified in the same way as the bromide, and it can be prepared from the iodide. A few milligrams dissolved in H<sub>2</sub>O can be purified by applying onto a Dowex-1 Cl<sup>-</sup> resin column (prepared by washing with N HCl followed by CO<sub>3</sub><sup>2-</sup>-free H<sub>2</sub>O until the pH is 5.8). After equilibration for 10 minutes, elution is started with CO<sub>3</sub><sup>2-</sup>-free distilled H<sub>2</sub>O, and 3ml fractions are collected and their OD values at λ 229nm are measured. The fractions with appreciable absorption are pooled and lyophilised at 0-5°. Note that at higher temperatures decomposition of the ester is appreciable; hydrolysis is appreciable at pH >10.5/20°. The residue is dried *in vacuo* over P<sub>2</sub>O<sub>5</sub>, checked for traces of iodine (add conc H<sub>2</sub>SO<sub>4</sub> and heat, violet vapours are released), and recrystallise it from propan-1-ol. [Gal & Roth *Clin Chim Acta* 2 316 1957, *Beilstein* 4 IV 1585.]

**S-Acetylthiocholine iodide** [1866-15-5] **M 289.2, m 203-204°, 204°, 204-205°**. Recrystallise the iodide from propan-1-ol (or *iso*-PrOH, or EtOH/Et<sub>2</sub>O) until almost colourless and dry it in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>. Its solubility in H<sub>2</sub>O is 1% w/v. A 0.075M (21.7mg/ml) solution in 0.1M phosphate buffer pH 8.0 is stable for 10-15 days if kept refrigerated. Store it away from light. It is commercially available as a 1% solution in H<sub>2</sub>O. [Ellman et al. *Biochemical Pharmacology* 7, 88 1961, IR: Hansen *Acta Chem Scand* 13 151 1959, 11 537 1957, *Clin Chim Acta* 2 316 1957, Ivin *Zh Obshch Khim* 22 267 1952, *Beilstein* 4 III 726, 4 IV 1585.]

**Actinomycin C (Cactinomycin)** [8052-16-2] **M ~1255**. (A commercial mixture of Actinomycin C<sub>1</sub> ~5%, C<sub>2</sub> ~30% and C<sub>3</sub> ~65% is available). *Actinomycin C<sub>1</sub> (native)* crystallises from EtOAc as red crystals, is soluble in CHCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub> and Me<sub>2</sub>CO, and has **m** 246-247°(dec), [α]<sub>D</sub><sup>20</sup> -328° (0.22, MeOH) and λ<sub>max</sub> 443nm (ε 25,000)

and 240nm ( $\epsilon$  34,000). *Actinomycin C<sub>2</sub> (native)* crystallises as red needles from EtOAc and has **m** 244-246°(dec),  $[\alpha]_{\text{D}}^{20}$  -325° (c 0.2, MeOH),  $\lambda_{\text{max}}$  443nm ( $\epsilon$  25,300) and 240nm ( $\epsilon$  33,400). *Actinomycin C<sub>3</sub> (native)* recrystallises from cyclohexane, or \*C<sub>6</sub>H<sub>6</sub>/MeOH/cyclohexane as red needles with **m** 238-241° (dec),  $[\alpha]_{\text{D}}^{20}$  -321° (c 0.2, MeOH),  $\lambda_{\text{max}}$  443nm ( $\epsilon$  25,000) and 240nm ( $\epsilon$  33,300). [Brockman & Lackner, *Chem Ber* **101** 1312 1968.] It is *light sensitive*. [Beilstein **27** III/IV 9642.]

**Actinomycin D (Dactinomycin)** [50-76-0] **M 1255.5, m 241-243°(dec)**,  $[\alpha]_{\text{D}}^{22}$ -296° (c 0.22, MeOH). It crystallises as bright red rhombic crystals from absolute EtOH or from MeOH/EtOH (1:3). It will also crystallise from EtOAc/cyclohexane (**m** 246-247° dec), CHCl<sub>3</sub>/petroleum ether (**m** 245-246° dec), and EtOAc/MeOH/\*C<sub>6</sub>H<sub>6</sub> (**m** 241-243° dec). Its solubility in MeCN is 1mg/ml.  $[\alpha]_{\text{D}}^{20}$  varies from -296° to -327° (c 0.2, MeOH).  $\lambda_{\text{max}}$  (MeOH) 445, 240nm (log  $\epsilon$  4.43, 4.49),  $\lambda_{\text{max}}$  (MeOH, 10N HCl, 1:1) 477nm (log  $\epsilon$  4.21) and  $\lambda_{\text{max}}$  (MeOH, 0.1N NaOH) 458, 344, 285 (log  $\epsilon$  3.05, 4.28, 4.13). It is **HIGHLY TOXIC**, light sensitive and anti-neoplastic. [Bullock & Johnson, *J Chem Soc* 3280 1957, Beilstein **27** III/IV 9642.]

**Adenosine-5'-diphosphate [adenosine-5'-pyrophosphate, ADP]** [58-64-0] **M 427.2,  $[\alpha]_{\text{D}}^{25}$  -25.7° (c 2, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> <2 (PO<sub>4</sub>H), pK<sub>2</sub><sup>25</sup> <2 (PO<sub>4</sub>H), pK<sub>3</sub><sup>25</sup> 3.95 (NH<sub>2</sub>), pK<sub>4</sub><sup>25</sup> 6.26 (PO<sub>4</sub>H)**. It is characterised by conversion to the *acridine salt* by addition of alcoholic acridine (1.1g in 50ml), filtering off the yellow salt and recrystallising from H<sub>2</sub>O. The salt has **m** 215°(dec), UV with  $\lambda_{\text{max}}$  at 259nm ( $\epsilon$  15,400) in H<sub>2</sub>O. [Baddiley & Todd *J Chem Soc* 648 1947, 582 1949, cf LePage *Biochemical Preparations* **1** 1 1949, Martell & Schwarzenbach *Helv Chim Acta* **39** 653 1956]. [Beilstein **26** III/IV 2369.]

**Adenosine-3'-monophosphoric acid hydrate [3'-adenylic acid, 3'-AMP]** [84-21-9] **M 347.3, m 197°(dec, as 2H<sub>2</sub>O), 210°(dec), m 210°(dec)**,  $[\alpha]_{546} -50°$  (c 0.5, 0.5M Na<sub>2</sub>HPO<sub>4</sub>), **pK<sub>1</sub><sup>25</sup> 3.65, pK<sub>2</sub><sup>25</sup> 6.05**. It crystallises from large volumes of H<sub>2</sub>O in needles as the *monohydrate*, but is not very soluble in boiling H<sub>2</sub>O. Under acidic conditions it forms an equilibrium mixture of 2' and 3' adenylic acids *via* the 2',3'-cyclic phosphate. When heated with 20% HCl, it gives a quantitative yield of furfural after 3 hours, unlike 5'-adenylic acid which only gives traces of furfural. The yellow *monoacridine salt* has **m** 175°(dec), and the *diacridine salt* has **m** 177° (225°)(dec). [Brown & Todd *J Chem Soc* 44 1952, Takaku et al. *Chem Pharm Bull Jpn* **21** 1844 1973, NMR: Ts' O et al. *Biochemistry* **8** 997 1969, Beilstein **26** III/IV 3607.]

**Adenosine-5'-monophosphoric acid monohydrate [5'-adenylic acid, 5'-AMP]** [18422-05-4] **M 365.2, m 178°, 196-200°, 200° (sintering at 181°)**,  $[\alpha]_{\text{D}}^{20}$  -47.5°,  $[\alpha]_{546} -56°$  (c 2, in 2% NaOH), -26.0° (c 2, 10% HCl), -38° (c 1, 0.5M Na<sub>2</sub>HPO<sub>4</sub>), **pK<sub>1</sub><sup>25</sup> 3.89, pK<sub>2</sub><sup>25</sup> 6.14, pK<sub>3</sub><sup>25</sup> 13.1**. The acid has been recrystallised from H<sub>2</sub>O (fine needles) and is freely soluble in boiling H<sub>2</sub>O. It crystallises also from H<sub>2</sub>O on addition of acetone. *Alternatively*, purify it by chromatography on Dowex 1 (in formate form), eluting with 0.25M formic acid. It is then adsorbed onto charcoal (which had been boiled for 15 minutes with M HCl, washed free of chloride and dried at 100°) and recovered by stirring three times with isoamyl alcohol/H<sub>2</sub>O (1:9 v/v). The aqueous layer from the combined extracts is evaporated to dryness under reduced pressure, and the product is crystallised twice from hot H<sub>2</sub>O. [Morrison & Doherty *Biochem J* **79** 433 1961]. It has UV with  $\lambda_{\text{max}}$  at 259nm ( $\epsilon$  15,400) in H<sub>2</sub>O at pH 7.0. [Alberty et al. *J Biol Chem* **193** 425 1951, Martell & Schwarzenbach *Helv Chim Acta* **39** 653 1956]. The *acridinium salt* has **m** 208° [Baddiley & Todd *J Chem Soc* 648 1947, Pettit *Synthetic Nucleotides*, van Nostrand-Reinhold, NY, **Vol 1** 252 1972, NMR: Sarma et al. *J Am Chem Soc* **96** 7337 1974, Norton et al. *J Am Chem Soc* **98** 1007 1976, IR of *diNa salt*: Miles *Biochem Biophys Acta* **27** 324 1958]. [Beilstein **26** III/IV 3615.]

**Adenosine 5'-[ $\beta$ -thio]diphosphate tri-lithium salt** [73536-95-5] **M 461.1**. Purify it by ion-exchange chromatography on DEAE-Sephadex A-25 using gradient elution with 0.1-0.5M triethylammonium bicarbonate. [Goody et al. *Biochem Biophys Acta* **276** 155 1972.]

**Adenosine 5'-[ $\alpha$ -thio]monophosphate di-lithium salt** [19341-57-2] **M 375.2**. Purify it as for the diNa salt [Murray & Atkinson *Biochemistry* **7** 4023 1968]. Dissolve 0.3g in dry MeOH (7ml) and 1M LiI (6ml) in dry Me<sub>2</sub>CO containing 1% of mercaptoethanol, and the Li salt is precipitated by adding Me<sub>2</sub>CO (75ml). The residue is washed with Me<sub>2</sub>CO (4 x 30ml) and dried at 55°/25mm. It has UV with  $\lambda_{\text{max}}$  at (HCl, pH 1.2) 257nm ( $\epsilon$  14,800); (0.015M NaOAc, pH 4.8) 259nm ( $\epsilon$  14,800); and (0.015M NH<sub>4</sub>OH, pH 10.1) 259nm ( $\epsilon$  15,300).

**Adenosine-5'-triphosphate (ATP)** [56-65-5] M 507.2,  $[\alpha]_{546}^{25} -35.5$  (c 1, 0.5 M Na<sub>2</sub>HPO<sub>4</sub>), pK<sub>1</sub><sup>25</sup> 4.00, pK<sub>2</sub><sup>25</sup> 6.48. ATP is purified by precipitating it as the barium salt on adding excess barium acetate solution to a 5% solution of ATP in water. The precipitate is filtered off, washed with distilled water, dissolved in 0.2M HNO<sub>3</sub> and again precipitated with barium acetate. The precipitate, after several washings with distilled water, is redissolved in 0.2M HNO<sub>3</sub>, and slightly more than an equivalent of 0.2M H<sub>2</sub>SO<sub>4</sub> is added to precipitate all the barium as BaSO<sub>4</sub> which is filtered off. The ATP is then precipitated by addition of a large excess of 95% ethanol. It is filtered off, washed several times with 100% EtOH and finally with dry diethyl ether. It is dried *in vacuo*. [Kashiwagi & Rabinovitch *J Phys Chem* 59 498 1955, *Beilstein* 26 III/IV 3654.]

**S-(5'-Adenosyl)-L-homocysteine** [979-92-0] M 384.4, m 202<sup>o</sup>(dec), 204<sup>o</sup>(dec), 205-207<sup>o</sup>(dec),  $[\alpha]_{D}^{25} +93^{\circ}$  (c 1, 0.2N HCl),  $[\alpha]_{D}^{23} +44^{\circ}$  (c 0.1, 0.05N HCl), (pK see SAM hydrochloride below). It has been recrystallised several times from aqueous EtOH or H<sub>2</sub>O to give small prisms and the UV has  $\lambda_{max}$  at 260nm in H<sub>2</sub>O. The *picrate* has m 170<sup>o</sup>(dec) from H<sub>2</sub>O. [Baddiley & Jamieson *J Chem Soc* 1085 1955, de la Haba & Cantoni *J Biol Chem* 234 603 1959, Borchardt et al. *J Org Chem* 41 565 1976, NMR: Follmann et al. *Eur J Biochem* 47 187 1974, *Beilstein* 26 III/IV 3676.]

**(-)-S-Adenosyl-L-methionine chloride (SAM hydrochloride)** [24346-00-7] M 439.9, pK<sub>Est(1)</sub>~ 2.13, pK<sub>Est(2)</sub>~ 4.12, pK<sub>Est(3)</sub>~ 9.28. Purify it by ion exchange on Amberlite IRC-150 and eluting with 0.1-4M HCl. [Stolowitz & Minch *J Am Chem Soc* 103 6015 1981.] It has been isolated as the tri-reineckate salt by adding 2 volumes of 1% solution of ammonium reineckate in 2% perchloric acid. The reineckate salt separates at once but is kept at 2<sup>o</sup> overnight. The salt is collected on a sintered glass funnel, washed with 0.5% of ammonium reineckate, dried (all operations at 2<sup>o</sup>) and stored at 2<sup>o</sup>. To obtain adenosylmethionine, the reineckate is dissolved in a small volume of methyl ethyl ketone (MEK) and centrifuged at room temperature to remove a small amount of solid. The clear dark red supernatant is extracted (in a separating funnel) with a slight excess of 0.1N H<sub>2</sub>SO<sub>4</sub>. The aqueous phase is re-extracted with fresh MEK until it is colourless. [Note that reineckates have UV absorption at 305nm ( $\epsilon$  15,000), and the optical density at 305nm is used to detect and estimate reineckate ions.] MEK is removed from the aqueous layer containing adenosylmethionine sulfate; the pH is adjusted to 5.6-6.0 and extracted with two volumes of Et<sub>2</sub>O. The *sulfate* is obtained by evaporating the aqueous layer *in vacuo*. The *hydrochloride* can be obtained in the same way but using HCl instead of H<sub>2</sub>SO<sub>4</sub>. SAM-HCl has a solubility of 10% in H<sub>2</sub>O. The salts are stable in the cold at pH 4-6 but decompose in alkaline media. [Cantoni *Biochemical Preparations* 5 58 1957.] The purity of SAM can be determined by paper chromatography [Cantoni *J Biol Chem* 204 403 1953, *Methods Enzymol* 3 601 1957], electrophoretic methods or enzymic analysis [Cantoni & Vignos *J Biol Chem* 209 647 1954]. [*Beilstein* 26 III/IV 3676.]

**L-Adrenaline** [*R*-(-)-epinephrine, L(-)-(3,4-dihydroxyphenyl)-2-methylaminoethanol] [51-43-4] M 183.2, m 210<sup>o</sup>(dec), 211<sup>o</sup>(dec), 211-212<sup>o</sup>(dec), 215<sup>o</sup>(dec),  $[\alpha]_{546}^{20} -61^{\circ}$  (c 5, 0.5M HCl),  $[\alpha]_{D}^{20} -52^{\circ}$  (c 2, 5% HCl), pK<sub>1</sub><sup>25</sup> 8.88 (8.75), pK<sub>2</sub><sup>25</sup> 9.90 (9.89), pK<sub>3</sub><sup>25</sup> 12.0 (~13). L-Adrenaline has been recrystallised from EtOH/AcOH/NH<sub>3</sub> [Jensen *J Am Chem Soc* 57 1765 1935]. It is sparingly soluble in H<sub>2</sub>O, readily in acidic or basic solutions but insoluble in aqueous NH<sub>3</sub>, alkali carbonate solutions, EtOH, CHCl<sub>3</sub>, Et<sub>2</sub>O or Me<sub>2</sub>CO. It has also been purified by dissolving in dilute aqueous acid, then precipitating it by adding dilute aqueous ammonia or alkali carbonates. It is readily oxidised in air and turns brown on exposure to light and air. (Epinephrine readily oxidises in neutral alkaline solution. This can be diminished if a little sulfite is added). Store it in the dark under N<sub>2</sub>. [Lewis *Br J Pharmacol Chemother* 9 488 1954]. The *hydrogen oxalate salt* has m 191-192<sup>o</sup>(dec, evacuated capillary) after recrystallisation from H<sub>2</sub>O or EtOH [Pickholz *J Chem Soc* 928 1945]. [*Beilstein* 13 H 830, 13 III/IV 2927.]

**Adrenolone hydrochloride** [3',4'-dihydroxy-2-methylaminoacetophenone hydrochloride] [62-13-5] M 217.7, m 244-249<sup>o</sup>(dec), 248<sup>o</sup>(dec), 256<sup>o</sup>(dec), pK<sub>25</sub> 5.5. It is purified by recrystallisation from EtOH or aqueous EtOH. [Gero *J Org Chem* 16 1222 1951, Kindler & Peschke *Arch Pharm* 269 581, 603 1931, *Beilstein* 14 IV 832.]

**Amethopterin (Methotrexate, 4-amino-4-deoxy-N<sup>10</sup>-methylpteroyl-L-glutamic acid)** [59-05-2] M 454.4, m 185-204<sup>o</sup>(dec),  $[\alpha]_{D}^{20} +19^{\circ}$  (c 2, 0.1N aqueous NaOH), pK<sub>1</sub> <0.5 (pyrimidine<sup>2+</sup>), pK<sub>2</sub> 2.5 (N5-Me<sup>+</sup>), pK<sub>3</sub> 3.49 ( $\alpha$ -CO<sub>2</sub>H), pK<sub>4</sub> 4.99 ( $\gamma$ -CO<sub>2</sub>H), pK<sub>5</sub> 5.50 (pyrimidine<sup>+</sup>). Most common impurities are 10-methylpteroyl-

glutamic acid, aminopterin and pteroylglutamic acid. Purify it by chromatography on Dowex-1 acetate, followed by filtration through a mixture of cellulose and charcoal. It has been recrystallised from aqueous HCl or by dissolution in the minimum volume of N NaOH and acidified until precipitation is complete, filter or *better* collect by centrifugation, wash with H<sub>2</sub>O (also by centrifugation) and dry at 100°/3mm. It has UV with  $\lambda_{\max}$  at 244 and 307nm ( $\epsilon$  17,300 and 19,700) in H<sub>2</sub>O at pH 1; 257, 302 and 370nm ( $\epsilon$  23,000, 22,000 and 7100) in H<sub>2</sub>O at pH 13. [Momle *Biochemical Preparations* **8** 20 1961, Seeger et al. *J Am Chem Soc* **71** 1753 1949.] It is a potent inhibitor of dihydrofolate reductase and is used in cancer chemotherapy. [Blakley *The Biochemistry of Folic Acid and Related Pteridines*, North-Holland Publ Co., Amsterdam, NY, pp157-163 1969, *Beilstein* **26** IV 3833.] It is **CARCINOGENIC; HANDLE WITH EXTREME CARE.**

**Aminopterin (4-amino-4-deoxypteroyl-L-glutamic acid)** [54-62-6] **M 440.4, m 231-235°(dec)**,  $[\alpha]_D^{20} +18^\circ$  (c 2, 0.1N aqueous NaOH), **pK<sub>1</sub> <0.5** (pyrimidine<sup>2+</sup>), **pK<sub>2</sub> 2.5** (N5-Me<sup>+</sup>), **pK<sub>3</sub> 3.49** ( $\alpha$ -CO<sub>2</sub>H), **pK<sub>4</sub> 4.65** ( $\gamma$ -CO<sub>2</sub>H), **pK<sub>5</sub> 5.50** (pyrimidine<sup>+</sup>). Purify aminopterin by recrystallisation from H<sub>2</sub>O. It has properties similar to those of methotrexate (above). It has UV at  $\lambda_{\max}$  244, 290 and 355nm ( $\epsilon$  18600, 21300 and 12000) in H<sub>2</sub>O at pH 1; 260, 284 and 370nm ( $\epsilon$  28500, 26400 and 8600) in H<sub>2</sub>O at pH 13. [Seeger et al. *J Am Chem Soc* **71** 1753 1949, Angier & Curran *J Am Chem Soc* **81** 2814 1959, Blakley *The Biochemistry of Folic Acid and Related Pteridines*, North-Holland Publ Co., Amsterdam, NY, pp 157-163 1969.] For small quantities, chromatograph it on DEAE cellulose with a linear gradient of ammonium bicarbonate pH 8 and increase the molarity from 0.1 to 0.4. Monitoring is by following the UV absorption of the fractions. For larger quantities, a near boiling solution of aminopterin (5g) in H<sub>2</sub>O (400ml) is slowly treated with small portions of MgO powder (~0.7g, calcined magnesia) with vigorous stirring until a small amount of MgO remained undissolved and the pH rises from 3-4 to 7-8. Charcoal (1g) is added to the hot solution and filtered immediately through a large sintered glass funnel of medium porosity and lined with a hot wet pad of Celite (~2-3 mm thick). The filtrate is cooled in ice, and the crystals of the Mg salt are collected by filtration and recrystallised from boiling H<sub>2</sub>O (200ml). The crystals are washed with EtOH and dried *in vacuo*. The Mg salt is redissolved in boiling H<sub>2</sub>O (200ml) and carefully acidified with vigorous agitation with AcOH (2ml). Pure aminopterin (3g) separates in fine yellow needles (dihydrate) which are easily filtered. The solid is washed with cold H<sub>2</sub>O, then Me<sub>2</sub>CO and dried *in vacuo*. If a trace of impurity is still present as shown by DEAE cellulose chromatography or TLC, repetition of the process will remove it; see UV above. [Loo *J Med Chem* **8** 139 1965, *Beilstein* **26** IV 3831.] **CARCINOGENIC.**

**3-Aminopyridine adenine dinucleotide** [21106-96-7] **M 635.4** (see **NAD for pK**). Purify it by ion-exchange chromatography as described [Fisher et al. *J Biol Chem* **248** 4293 1973, Anderson & Fisher *Methods Enzymol* **66** 81 1980].

**Anion exchange resins.** These should be conditioned before use by successive washing with water, EtOH and water, and taken through two OH<sup>-</sup>—H<sup>+</sup>—OH<sup>-</sup> cycles by successive treatment with N NaOH, water, N HCl, water and N NaOH, then washed with water until neutral to give the OH<sup>-</sup> form. (See commercial catalogues on ion-exchange resins.)

**Aniracetam (1-[4-methoxybenzyl]-2-pyrrolidinone)** [72432-10-1] **M 219.1, m 121-122°**. Purify aniracetam by recrystallisation from EtOH. It is a nootropic (Alzheimer) drug. [Gouliarov & Senning *Brain Research Rev* **19** 180 1994.]

**Apocodeine** [641-36-1] **M 281.3, m 124°, pK<sub>Est(1)</sub> ~ 7.0, pK<sub>Est(2)</sub> ~ 8.2**. Crystallise apocodeine from absolute EtOH by boiling and allowing to stand at 0° to give the alcoholate, **m 104.5-106.5°**; EtOH is lost slowly at 25°/2mm but readily at 78°/2mm, and the anhydrous base has **m 122.5-124.5°**. It is soluble in Et<sub>2</sub>O. [Folkers *J Am Chem Soc* **58** 1814 1936.] It has also been crystallised from MeOH or MeOH with a little CH<sub>2</sub>Cl<sub>2</sub> as a waxy solid and dried at 80°/2mm. It softens above 100° before melting and is sensitive to air and light. The *hydrochloride*, which is formed from the base in EtOH containing an equivalent amount of HCl followed by addition of Et<sub>2</sub>O, is recrystallised from 95% EtOH and adding Et<sub>2</sub>O until crystals separate. It melts at 260-263° (dec, with softening at 140°) and has  $[\alpha]_D^{20} -41.3^\circ$  to  $-43.3^\circ$  (c 0.81, H<sub>2</sub>O). [Neumyer et al. *J Med Chem* **16** 1223 1973.]

**R-Apomorphine** [58-00-4] M 267.3, m 195°(dec), pK<sub>1</sub><sup>15</sup> 7.20 (NH<sub>2</sub>), pK<sub>2</sub><sup>15</sup> 8.91 (phenolic OH). Crystallise R-apomorphine from CHCl<sub>3</sub> and a little petroleum ether, also from Et<sub>2</sub>O with 1 mol of Et<sub>2</sub>O which it loses at 100°. It sublimes in a high vacuum. It is white but turns green in moist air or in alkaline solution. Its UV has λ<sub>max</sub> at 336, 399 (98% EtOH). The *di-O-methylether* is an oil b 175°/high vacuum, whose *picrate* crystallises from MeOH and has m 140° (dec). The *di-O-acetate* crystallises from EtOAc/petroleum ether with m 127-128°, [α]<sub>D</sub><sup>25</sup> -88° (c 1, 0.1 N HCl). The *di-O-benzoyl* derivative has m 156-158° (from EtOH) and [α]<sub>D</sub><sup>18</sup> +43.44° (c 3.3, CHCl<sub>3</sub>). [Pachorr et al. *Chem Ber* 35 4377 1902, *Beilstein* 21 H 246.] **NARCOTIC.**

**R-Apomorphine hydrochloride** [41372-20-7] M 312.8, m 285-287°(dec), [α]<sub>D</sub><sup>20</sup> -48° (c 1, H<sub>2</sub>O). Crystallise the salt from H<sub>2</sub>O (*hemihydrate*) and from EtOH. Crystals turn green on exposure to light. (see previous entry). **NARCOTIC.**

**Aureomycin (7-chlorotetracycline)** [57-62-5] M 478.5, m 172-174°(dec), [α]<sub>D</sub><sup>23</sup> -275° (MeOH), pK<sub>1</sub> 3.3, pK<sub>2</sub> 7.44, pK<sub>3</sub> 9.27. Aureomycin is dehydrated by azeotropic distillation of its solution with toluene. On cooling, the anhydrous material crystallises out and is recrystallised from \*C<sub>6</sub>H<sub>6</sub>, then dried under vacuum at 100° over paraffin wax. (If it is crystallised from MeOH, it contains MeOH which is not removed on drying.) [Stephens et al. *J Am Chem Soc* 76 3568 1954, Laskin & Chan *Biochem Biophys Res Commun* 14 137 1964]. [*Beilstein* 14 IV 2631.]

**Aureomycin hydrochloride (7-chlorotetracycline hydrochloride)** [64-72-2] M 514.0, m 234-236°(dec), [α]<sub>D</sub><sup>25</sup> -23.5° (H<sub>2</sub>O). Purify the salt by dissolving 1g rapidly in 20ml of hot water, cooling rapidly to 40°, treating with 0.1ml of 2M HCl, and chilling in an ice-bath. The process is repeated twice. It is also recrystallised from Me<sub>2</sub>NCHO/Me<sub>2</sub>CO. [Stephens et al. *J Am Chem Soc* 76 3568 1954, UV: McCormick et al. *J Am Chem Soc* 79 2849 1957, *Beilstein* 14 IV 2631.]

**Bacitracin (Altracin, Topitracin)** [1405-87-4] M 1422.7, [α]<sub>D</sub><sup>23</sup> +5° (H<sub>2</sub>O). Bacitracin has been purified by carrier displacement using *n*-heptanol, *n*-octanol and *n*-nonanol as carriers and 50% EtOH in 0.1 N HCl. The pure material gives one spot with R<sub>F</sub> ~0.5 on paper chromatography using AcOH:*n*-BuOH:H<sub>2</sub>O (4:1:5). [Porath *Acta Chem Scand* 6 1237 1952.] It has also been purified by ion-exchange chromatography. It is a white powder soluble in H<sub>2</sub>O and EtOH but insoluble in Et<sub>2</sub>O, CHCl<sub>3</sub> and Me<sub>2</sub>CO. It is stable in acidic solution but unstable in base. It is a strong antibacterial. [Abraham & Bewton *Biochem J* 47 257 1950, Synthesis: Munekata et al. *Bull Chem Soc Jpn* 46 3187, 3835 1973, *Beilstein* 27 III/IV 5746.]

**N<sup>6</sup>-Benzyladenosine** [4294-16-0] M 357.4, m 177-179°, 185-187°, [α]<sub>D</sub><sup>25</sup> -68.6° (c 0.6, EtOH)(see pK of adenosine). Purify it by recrystallisation from EtOH. It has UV with λ<sub>max</sub> at 266nm (aqueous EtOH/HCl) and 269 nm (aqueous EtOH/NaOH). [Kissman & Weiss *J Org Chem* 21 1053 1956, *Beilstein* 26 III/IV 3682.]

**N-Benzylpenicillin sodium salt (penicillin G Na salt)** [69-57-8] M 356.37, m 215° (charring and dec), 225° (dec), [α]<sub>D</sub><sup>20</sup> +269° (c 0.7, MeOH), [α]<sub>D</sub><sup>25</sup> +305° (c 1, H<sub>2</sub>O), pK<sup>25</sup> 2.76 (4.84 in 80% aqueous EtOH)(for free acid). Purify the salt by dissolving it in a small volume of MeOH (in which it is more soluble than EtOH) and treating gradually with ~5 volumes of EtOAc. This gives an almost colourless crystalline solid (rosettes of clear-cut needles) and recrystallising twice more if slightly yellow in colour. The salt has also been conveniently recrystallised from the minimum volume of 90% Me<sub>2</sub>CO and adding an excess of absolute Me<sub>2</sub>CO. A similar procedure can be used with wet *n*-BuOH. If yellow in colour, then dissolve (~3.8g) in the minimum volume of H<sub>2</sub>O (3ml), add *n*-BuOH and filter through a bed of charcoal. The salt forms long white needles on standing in a refrigerator overnight. More crystals can be obtained on concentrating the mother liquors *in vacuo* at 40°. <sup>3</sup>A further recrystallisation (without charcoal) yields practically pure salt. A good preparation has ~600 Units/mg. The presence of H<sub>2</sub>O in the solvents increases the solubility considerably. The solubility in mg/100ml at 0° is 6.0 (Me<sub>2</sub>CO), 15.0 (Me<sub>2</sub>CO/0.5% H<sub>2</sub>O), 31.0 (Me<sub>2</sub>CO/1.0% H<sub>2</sub>O), 2.4 (methyl ethyl ketone), 81.0 (*n*-butanol) and 15.0 (dioxane at 14°). *Alternatively*, it is dissolved in H<sub>2</sub>O (solubility is ~10%), filtered if necessary and precipitated by addition of EtOH and dried in a vacuum over P<sub>2</sub>O<sub>5</sub>. A sample can be kept for 24 hours at 100° without loss of physiological activity. It also crystallises from MeOH/EtOAc. [IR: Barnes et al. *Anal Chem* 19 620 1947, *The Chemistry of Penicillin* (Clarke, Johnson and Robinson eds.) Princeton University Press,

Princeton NJ, Chapter V 85 1949, *Beilstein* 27 III/IV 5861.]

Other salts, e.g. the **potassium salt** (M 372.5 [113-58-4]) can be prepared from the Na salt by dissolving it (147mg) ice-cold in H<sub>2</sub>O acidified to pH 2, extracting with Et<sub>2</sub>O (~50ml), washing once with H<sub>2</sub>O, and extracting with 2ml portions of 0.3% KHCO<sub>3</sub> until the pH of the extract rises to ~6.5 (~7 extractions). The combined aqueous extracts are lyophilised, and the white residue is dissolved in *n*-BuOH (1ml, absolute) with the addition of enough H<sub>2</sub>O to effect solution. Remove insoluble material by centrifugation and add absolute *n*-BuOH to the supernatant. Crystals should separate on scratching, and after 2.5 hours in a refrigerator they are collected, washed with absolute *n*-BuOH and EtOAc and dried (yield 51.4mg). It also crystallises from aqueous Me<sub>2</sub>CO. The *potassium salt* has **m** 214-217° (dec) (block preincubated at 200°; heating rate of 3°/min) and  $[\alpha]_{\text{D}}^{22} +285^{\circ}$  (c 0.748, H<sub>2</sub>O). [*Beilstein* 27 III/IV 5861.]

The **free acid (penicillin G)** (M 334.5, [61-33-6]) has **m** 186-187° (MeOH/Me<sub>2</sub>CO), and **m** 190-191° (H<sub>2</sub>O)  $[\alpha]_{\text{D}}^{25} +282^{\circ}$  (EtOH). [Review: Sheehan & Henery-Logan *J Am Chem Soc* 84 2983 1962, Chain et al. *Antibiotics* (Oxford University Press) 2 1949, and Cook *Quarterly Reviews (Chemical Society)* 2 46 1948.]

**Bergapten (5-methoxypsoralen)** [484-20-8] **M 216.2, m 190-193°, 191-193°**. Crystallise it from EtOH or aqueous MeOH, and it sublimes *in vacuo*. Its properties are similar to those of its 9-methoxy isomer (xanthotoxin, see below). It is slightly soluble in \*C<sub>6</sub>H<sub>6</sub>, CHCl<sub>3</sub> and AcOH but insoluble in H<sub>2</sub>O. Its fluorescence has  $\lambda_{\text{ex}}$  at 352nm with  $\lambda_{\text{em}}$  at 480nm. It is a DNA intercalator and **possible carcinogen**. [Howell & Robertson *J Chem Soc* 293 1937, Boyer et al. *Biochemistry* 27 3011 1988, *Beilstein* 19/6 V 4.]

**(+)-Bicuculine** [*R*-6(5,6,7,8-tetrahydro-6-methyl-1,3-dioxolo[4,5-*g*]isoquinolin-5-yl)-furo-[3,4-*c*]-1,3-benzodioxolo-8(6*H*)-one] [485-49-4] **M 367.4, m 177°, 193-195°, 193-197°, 215°,  $[\alpha]_{\text{D}}^{20} +126^{\circ}$  (c 1, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{20} +159^{\circ}$  (c 1, CHCl<sub>3</sub>), pK<sup>25</sup> 4.84**. It crystallises from CHCl<sub>3</sub>/MeOH as plates. The crystals melt at 177°, then solidify and re-melt at 193-195° [Manske *Canad J Research* 21B 13 1943]. It is soluble in CHCl<sub>3</sub>, \*C<sub>6</sub>H<sub>6</sub>, EtOAc but sparingly soluble in EtOH, MeOH and Et<sub>2</sub>O. [Stereochem: Blaha et al. *Col Czech Chem Commun* 29 2328 1964, Snatzke et al. *Tetrahedron* 25 5059 1969, Pharmacol: Curtis et al. *Nature* 266 1222 1970, *Beilstein* 27 III/IV 1900].

**L-erythro-Biopterin** (2-amino-4-hydroxy-6-[(1*R*,2*S*)-1,2-dihydroxypropyl]pteridine) [22150-76-1] **M 237.2, m >300°(dec),  $[\alpha]_{\text{D}}^{20} -80^{\circ}$ ,  $[\alpha]_{\text{D}}^{20} -65^{\circ}$  (c 2.0, M HCl), pK<sup>25</sup> 2.23(2.45), pK<sup>25</sup> 7.89(8.05)**. Purify L-erythro-biopterin by chromatography on Florisil, which was washed thoroughly with 2M HCl, and elute with 2M HCl. The fractions with UV-fluorescence are pooled, evaporated *in vacuo* and the residue recrystallised. Biopterin is best recrystallised (90% recovery) by dissolving in 1% aqueous NH<sub>3</sub> (*ca* 100 parts), and adding this solution dropwise to an equal volume of M aqueous formic acid at 100° and allowing to cool at 4° overnight. It is dried at 20° to 50°/0.1mm in the presence of P<sub>2</sub>O<sub>5</sub>. [Schircks et al. *Helv Chim Acta* 60 211 1977, Armarego et al. *Aust J Chem* 35 785 1982.] It also crystallises from *ca* 50 parts of water or 100 parts of hot 3M aqueous HCl by adding hot 3M aqueous NH<sub>3</sub> and cooling. It has UV with  $\lambda_{\text{max}}$  at 212, 248 and 321nm (log  $\epsilon$  4.21, 4.09 and 3.94) in H<sub>2</sub>O at pH 0.0; 223infl, 235.5, 274.5 and 345nm (log  $\epsilon$  4.07infl, 4.10, 4.18 and 3.82) in H<sub>2</sub>O at pH 5.0; 221.5, 254.5 and 364nm (log  $\epsilon$  3.92, 4.38 and 3.84) in H<sub>2</sub>O at pH 10.0. [Sugimoto & Matsuura *Bull Chem Soc Jpn* 48 3767 1875, *Beilstein* 26 III/IV 4032.]

**D-(+)-Biotin (vitamin H, hexahydro-2-oxo-1*H*-thieno[3,4-*d*]imidazole-4-pentanoic acid)** [58-85-5] **M 244.3, m 229-231°, 230.2°(dec), 230-231°, 232-234°(dec),  $[\alpha]_{\text{D}}^{20} +108^{\circ}$ ,  $[\alpha]_{\text{D}}^{20} +91.3^{\circ}$  (c 1, 0.1N NaOH), pK<sub>Est</sub> ~ 4.8**. D-(+)-Biotin crystallises from hot water in fine long needles with a solubility of 22 mg/100ml at 25°. Its solubility in 95% EtOH is 80 mg/100 ml at 25°. Its isoelectric point is at pH 3.5. Store solid and solutions under sterile conditions because it is susceptible to mould growth. [Confalone *J Am Chem Soc* 97 5936 1975, Wolf et al. *J Am Chem Soc* 67 2100 1945, Synthesis: Ohuri & Emoto *Tetrahedron Lett* 2765 1975, Harris et al. *J Am Chem Soc* 66 1756 1944.] The (+)-*methyl ester* has **m** 166-167° (from MeOH/Et<sub>2</sub>O),  $[\alpha]_{\text{D}}^{22} +57^{\circ}$  (c 1, CHCl<sub>3</sub>) [du Vigneaud et al. *J Biol Chem* 140 643, 763 1941]; the (+)-*S-oxide* has **m** 200-203°,  $[\alpha]_{\text{D}}^{20} +130^{\circ}$  (c 1.2, 0.1N NaOH) [Melville *J Biol Chem* 208 495 1954]; the *SS-dioxide* has **m** 274-275°(dec, 268-270°), and the *SS-dioxide methyl ester* has **m** 239-241° (from MeOH/Et<sub>2</sub>O) [Hofmann et al. *J Biol Chem* 141 207, 213 1941]. [*Beilstein* 27 III/IV 7979.]

**D-(+)-Biotin hydrazide** [66640-86-6] M 258.4, m 238-240°, 245-247°,  $[\alpha]_D^{20} +66^\circ$  (c 1, Me<sub>2</sub>NCHO). Wash the hydrazide with H<sub>2</sub>O, dry it, wash it with MeOH then Et<sub>2</sub>O, and dry. Recrystallise it from hot H<sub>2</sub>O (clusters of prisms) [Hofmann et al. *J Biol Chem* **144** 513 1942]. [Beilstein **27** III/IV 7980.]

**D-(+)-Biotin N-hydroxysuccinimide ester (+-biotin N-succinimidyl ester)** [35013-72-0] M 342.4, m 210°, 212-214°,  $[\alpha]_D^{20} +53^\circ$  (c 1, Me<sub>2</sub>NCHO). Recrystallise the ester from refluxing isoPrOH and dry it in a vacuum over P<sub>2</sub>O<sub>5</sub> + KOH. [Jasiewicz et al. *Exp Cell Biol* **100** 213 1976.]

**D-(+)-Biotin 4-nitrophenyl ester** [33755-53-2] M 365.4, m 160-163°, 163-165°,  $[\alpha]_D^{25} +47^\circ$  (c 2, Me<sub>2</sub>NCHO containing 1% AcOH). The ester has been recrystallised by dissolving 2g in 95% EtOH (30ml), heated to dissolve, then cooled in an ice-water bath. The crystals are collected, washed with ice-cold 95% EtOH (5ml) and dried over P<sub>2</sub>O<sub>5</sub>. The R<sub>F</sub> on silica plates (CHCl<sub>3</sub>/MeOH 19:1) is 0.19 [Bodanszky & Fagan *J Am Chem Soc* **99** 235 1977].

**N-(+)-Biotinyl-4-aminobenzoic acid** [6929-40-4] M 363.4, m 295-297°, 295-300°,  $[\alpha]_D^{23} +56.55^\circ$  (c 0.5, 0.1N NaOH), pK<sub>Est</sub> ~4.0. Dissolve the acid in NaHCO<sub>3</sub> solution, cool and precipitate it by adding N HCl. Collect the solid, dry it at 100° and recrystallise it from MeOH. Note that it is hydrolysed by aqueous 3M, 1M and 0.2M HCl at 120°, but can be stored in 5% aqueous NaHCO<sub>3</sub> at -20° without appreciable hydrolysis [Knappe et al. *Biochem Zeitschrift* **338** 599 1963, Wolf et al. *J Am Chem Soc* **73** 4142 1951, Bayer & Wilchek *Methods Enzymol* **26** 1 1980]. [Beilstein **27** III/IV 7984.]

**N-Biotinyl-6-aminocaproic N-succinimidyl ester** [72040-63-2] M 454.5, m 149-152°. Dissolve ~400mg of the ester in dry propan-2-ol (~25ml) with gentle heating. Reduce the volume to ~10ml by gentle boiling and allow the solution to cool. Decant the supernatant carefully from the white crystals, dry the crystals in a vacuum over P<sub>2</sub>O<sub>5</sub> at 60° overnight. This material gives one spot on TLC. [Costello et al. *Clin Chem* **25** 1572 1979, Kincaid et al. *Methods Enzymol* **159** 619 1988.]

**N-(+)-Biotinyl-6-aminocaproyl hydrazide (biotin-6-aminohexanoic hydrazide)** [109276-34-8] M 371.5, m 189-191°, 210°,  $[\alpha]_D^{20} +23^\circ$  (c 1, Me<sub>2</sub>NCHO). Suspend the hydrazide in ice-water (100mg/ml), stand overnight at 4°, filter and dry the solid in a vacuum. Recrystallise it from isoPrOH. R<sub>F</sub> is 0.26 on SiO<sub>2</sub> plate using CHCl<sub>3</sub>/MeOH (7:3) as eluent. [O'Shannessy et al. *Anal Biochem* **163** 204 1987.]

**N-(+)-Biotinyl-L-lysine (Biocytin)** [576-19-2] M 372.5, m 228.5°, 228-230° (dec), 241-243°, 245-252° (dec, sintering at 227°),  $[\alpha]_D^{25} +53^\circ$  (c 1.05, 0.1 N NaOH). Recrystallise biocytin rapidly from dilute MeOH or Me<sub>2</sub>CO. It can also be recrystallised from H<sub>2</sub>O by slow evaporation or by dissolving in the minimum volume of H<sub>2</sub>O and adding Me<sub>2</sub>CO until solid separates. It is freely soluble in H<sub>2</sub>O and AcOH but insoluble in Me<sub>2</sub>CO. [Wolf et al. *J Am Chem Soc* **74** 2002 1952, **72** 1048 1950.] It has been purified by chromatography on superfiltrol-Celite, Al<sub>2</sub>O<sub>3</sub> and by countercurrent distribution and then recrystallised [IR: Peck et al. *J Am Chem Soc* **74** 1991 1952]. The *hydrochloride* recrystallises from aqueous Me<sub>2</sub>CO/HCl and has m 227° (dec). [Beilstein **27** III/IV 7984.]

**Brefeldin A [1-R-2c,15c-dihydroxy-7t-methyl-(1r,13t)-6-oxa-bicyclo[11.3.0]hexadeca-3t,11t-dien-5-one, Decumbin]** [20350-15-6] M 280.4, m 200-202°, 204°, 204-205°,  $[\alpha]_D^{22} +95^\circ$  (c 0.81, MeOH). Brefeldin A was isolated from *Penicillium brefeldianum* and recrystallised from aqueous MeOH/EtOAc or MeOH. Its solubility in H<sub>2</sub>O is 0.6mg/ml, 10mg/ml in MeOH and 24.9mg/ml in EtOH. The *O-acetate* recrystallises from Et<sub>2</sub>O/pentane and has m 130-131°,  $[\alpha]_D^{22} +17^\circ$  (c 0.95, MeOH). [Sigg *Helv Chim Acta* **47** 1401 1964, UV and IR: Härrri et al. *Helv Chim Acta* **46** 1235 1963, total synthesis: Kitahara et al. *Tetrahedron* 3021 1979, X-ray: Weber et al. *Helv Chim Acta* **54** 2763 1971, Beilstein **18** III/IV 1220.]

**5-Bromo-2'-deoxyuridine** [59-14-3] M 307.1, m 193-197°(dec), 217-218°,  $[\alpha]_D^{25} -4.1^\circ$  (c 0.1, H<sub>2</sub>O), pK<sup>25</sup> ~8.1. Recrystallise the uridine from EtOH or 96% EtOH. It has UV with  $\lambda_{max}$  at 279 nm at pH 7.0, and 279 nm (log  $\epsilon$  3.95) at pH 1.9. Its R<sub>F</sub> values are 0.49, 0.46 and 0.53 in *n*-BuOH/AcOH/H<sub>2</sub>O (4:1:1), *n*-BuOH/EtOH/H<sub>2</sub>O (40:11:19) and *i*-PrOH-25% aqueous NH<sub>3</sub>-H<sub>2</sub>O (7:1:1), respectively. [Nature **209** 230 1966, Prystās & Sorm *Col Czech Chem Comm* **29** 2956 1964, Beilstein **24** III/IV 1234.]

**5-Bromouridine** [957-75-5] **M 323.1, m 215-217°**, 217-218°,  $[\alpha]_D^{22}$  -24.1° (c 2, H<sub>2</sub>O), **pK<sup>25</sup> 8.1**. Recrystallise it from 96% EtOH. It has UV with  $\lambda_{\max}$  at 279nm (log  $\epsilon$  3.95) in H<sub>2</sub>O pH 1.9.  $R_F$  in *n*-BuOH/AcOH/H<sub>2</sub>O (4:4:1) is 0.49; in *n*-BuOH/EtOH/H<sub>2</sub>O (40:11:9) it is 0.46 and in isoPrOH/25%NH<sub>3</sub>/H<sub>2</sub>O (7:1:2) it is 0.53 using Whatman No 1 paper. [Prystās & Sorm *Col Czech Chem Commmun* **29** 2956 1964, *Beilstein* **31** H 24.]

**Brucine** [357-57-3 (anhydrous), 5892-11-5 (4H<sub>2</sub>O)] **M 394.5 (anhydrous), 430.5 (2H<sub>2</sub>O), 466.5 (4H<sub>2</sub>O), m 178-179°**,  $[\alpha]_{546}^{20}$  -149.9° (anhydrous; c 1, in CHCl<sub>3</sub>),  $[\alpha]_D^{20}$  -119° (anhydrous; c 2, in CHCl<sub>3</sub>), **pK<sub>1</sub><sup>15</sup> 2.50, pK<sub>2</sub><sup>15</sup> 8.16 (pK<sub>2</sub><sup>25</sup> 8.28)**. Crystallise brucine once from water or aqueous Me<sub>2</sub>CO (as the *tetrahydrate*), then suspend it in CHCl<sub>3</sub> and shake with anhydrous Na<sub>2</sub>SO<sub>4</sub> (to dehydrate the brucine, which then dissolves). Precipitate it by pouring the solution into a large bulk of dry petroleum ether (b 40-60°), filter and heat to 120° in a high vacuum [Turner *J Chem Soc* 842 1951]. The *tetrahydrate* crystallises from a mixture of EtOH and H<sub>2</sub>O as colourless elongated needles [Eeles *Acta Cryst* **6** 809 1953, *Beilstein* **27** III/IV 7875.] **VERY POISONOUS.**

**Brucine sulfate (hydrate)** [4845-99-2] **M 887.0, m ~180°(dec)**,  $[\alpha]_{546}^{20}$  -36° (dry; c 1, H<sub>2</sub>O). The *heptahydrate* crystallises from water as colourless laths. [Eeles *Acta Cryst* **6** 809 1953, *Beilstein* **27** III/IV 7875.]

**Butyryl choline iodide [(2-butyryloxyethyl)trimethyl ammonium iodide]** [2494-56-6] **M 301.7, m 85-89°, 87°, 93-94°**. Recrystallise the iodide from isoPrOH or Et<sub>2</sub>O. [Tammelin *Acta Chem Scand* **10** 145 1956.] The *perchlorate* has **m 72°** (from isoPrOH). [Aldridge *Biochem J* **53** 62 1953, *Beilstein* **4** IV 1448.]

**S-Butyryl thiocholine iodide [(2-butyrylmercaptoethyl)trimethyl ammonium iodide]** [1866-16-6] **M 317.2, m 173°, 173-176°**. Recrystallise S-butyryl thiocholine iodide from propan-1-ol and dry it *in vacuo*; store it in the dark under N<sub>2</sub>. The *bromide* has **m 150°** (from Me<sub>2</sub>CO) or **m 140-143°** (from butan-1-ol). [Gillis *Chem Ind (London)* 111 1957, Hansen *Acta Chem Scand* **11** 537 1957, *Beilstein* **4** IV 1586.]

**Carminic acid** (7- $\alpha$ -D-glucopyranosyl-9,10-dihydro-3,5,6,8-tetrahydroxy-1-methyl-9,10-dioxo-2-anthracene carboxylic acid, **Neutral Red 4: CI 75470**) [1260-17-9] **M 492.4, m 120°(dec)**,  $[\alpha]_{654}^{15}$  +51.6° (H<sub>2</sub>O), (several phenolic pKs). Carminic acid forms red prisms from EtOH. It gives a red colour in Ac<sub>2</sub>O and yellow to violet in acidic solution. Its UV has  $\lambda_{\max}$  (H<sub>2</sub>O) at 500nm ( $\epsilon$  6,800); (0.02N HCl) 490-500nm ( $\epsilon$  5,800) and (0.0001N NaOH) 540nm ( $\epsilon$  3,450). Its IR has  $\nu_{\max}$  at (Nujol) 1708s, 1693s, 1677m, 1648m, 1632m, 1606s, 1566s, 1509 cm<sup>-1</sup>. Periodate oxidation is complete after 4 hours at 0° with the consumption of 6.2 mols. The *tetra-O-methyl carminate* has **m 186-188°** (yellow needles from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether). [IR: Ali & Haynes *J Chem Soc* 1033 1959, Bhatia & Venkataraman *Indian J Chem* **3** (2) 92 1965, Synthesis: Davis & Smith *Biochemical Preparations* **4** 38 1955, *Beilstein* **18** III/IV 6697.]

**Cation exchange resins.** These should be conditioned before use by successive washing with water, EtOH and water, and taken through two H<sup>+</sup>-OH<sup>-</sup>-H<sup>+</sup> cycles by successive treatment with M HCl, water, M NaOH, water and M HCl, then washed with water until neutral to give the H<sup>+</sup> form. (See commercial catalogues on ion exchange resins).

**Cephalosporin C potassium salt** [28240-09-7] **M 453.5,  $[\alpha]_D^{20}$  +103° (H<sub>2</sub>O), pK<sub>1</sub> <2.6, pK<sub>2</sub> 3.1, pK<sub>3</sub> 9.8**. Purify the salt by dissolving it in the minimum volume of H<sub>2</sub>O (filter) and adding EtOH until separation of solid is complete. A solution is stable in the pH range 2.5-8. It has UV with  $\lambda_{\max}$  at 260nm (log  $\epsilon$  3.95) in H<sub>2</sub>O. The **Ba salt** has  $[\alpha]_D^{20}$  +80° (c 0.57, H<sub>2</sub>O) [Woodward et al. *J Am Chem Soc* **88** 852 1966, Abraham & Newton *Biochem J* **79** 377 1961, Hodgkin & Maslen *Biochem J* **79** 402 1961; see also *Quart Reviews Chem Soc* London **21** 231 1967]. [*Beilstein* **27** III/IV 5902.]

**Chlorambucil** [4-{bis(2-chloroethyl)amino}-phenylbutyric acid] [305-03-3] **M 304.2, m 64-66°, pK<sub>1</sub> 5.8 (6.0 at 66°, 50% aqueous Me<sub>2</sub>CO), pK<sub>2</sub> 8.0**. Chlorambucil is recrystallised from petroleum ether (flat needles) and has a solubility at 20° of 66% in EtOH, 40% in CHCl<sub>3</sub>, 50% in Me<sub>2</sub>CO but is insoluble in H<sub>2</sub>O [Everett et al. *J Chem Soc* 2386 1953]. [*Beilstein* **14** IV 1715.] **CARCINOGEN.**

**Chloramphenicol** [Amphicol, 1R,2R(-)-2-{2,2-dichloroacetyl-amino}-1-{4-nitrophenyl}-propan-1,3-diol]

[56-75-7] **M 323.1, m 149-151°, 150-151°, 151-152°,  $[\alpha]_D^{20} +20.5^\circ$  (c 3, EtOH),  $[\alpha]_D^{25} -25.5^\circ$  (EtOAc).** Purify chloramphenicol by recrystallisation from H<sub>2</sub>O (solubility is 2.5mg/ml at 25°) or ethylene dichloride as needles or long plates, and by sublimation at high vacuum. It has  $A_{1\text{cm}}^{1\%}$  298 at  $\lambda_{\text{max}}$  278nm, and it is slightly soluble in H<sub>2</sub>O (0.25%) and propylene glycol (1.50%) at 25° but is freely soluble in MeOH, EtOH, BuOH, EtOAc and Me<sub>2</sub>CO. [Relstock et al. *J Am Chem Soc* **71** 2458 1949, Confroulis et al. *J Am Chem Soc* **71** 2463 1949, Long & Troutman *J Am Chem Soc* **71** 2469, 2473 1949, Ehrhart et al. *Chem Ber* **90** 2088 1957, *Beilstein* **13** IV 2742.]

**Chloramphenicol palmitate** [530-43-8] **M 561.5, m 90°,  $[\alpha]_D^{26} +24.6^\circ$  (c 5, EtOH).** The palmitate crystallises from \*benzene or xylene with **m** 105-106° and  $[\alpha]_D^{21} -39.5^\circ$  (c 2, Et<sub>2</sub>O),  $\lambda_{\text{max}}$  267.3nm. [Edgerton et al. *J Am Chem Soc* **77** 27 1955, *Beilstein* **13** IV 2753.]

**2-Chloroadenosine** [146-77-0] **M 301.7, m 145-146°(dec), 147-149°(dec),  $\text{pK}_{\text{Est}(1)} \sim 0.5$ ,  $\text{pK}_{\text{Est}(2)} \sim 7.6$ .** Purify 2-chloroadenosine by recrystallisation from H<sub>2</sub>O (~1% in cold), and it has UV with  $\lambda_{\text{max}}$  at 264 nm (pH 1 and 7) and 265 nm (pH 13) in H<sub>2</sub>O. [Brown & Weliky *J Org Chem* **23** 125 1958, Schaeffer & Thomas *J Am Chem Soc* **80** 3738 1958, IR: Davoll & Lewy *J Am Chem Soc* **74** 1563 1952, *Beilstein* **26** III/IV 3725.]

**Chlorophyll a** [479-61-8] **M 983.5, m 117-120°, 150-153°, 178-180° (sinters at ~150°),  $[\alpha]_D^{20} -262^\circ$  (Me<sub>2</sub>CO).** It forms green crystals from Me<sub>2</sub>CO, Et<sub>2</sub>O/H<sub>2</sub>O, Et<sub>2</sub>O/hexane/H<sub>2</sub>O or Et<sub>2</sub>O/pentane /H<sub>2</sub>O. It is sparingly soluble in MeOH and insoluble in petroleum ether. In alkaline solution it gives a blue-green colour with deep red fluorescence. A very crude chlorophyll mixture has been purified by chromatography on low melting polyethylene (MI 0.044; "Dow" melting index MI <2) and developed with 70% aqueous Me<sub>2</sub>CO. The order of effluent from the bottom of the column is: xanthophylls, chlorophyll *b*, chlorophyll *a*, phaeophytins and carotenes. A mixture of chlorophylls *a* and *b* is best separated by chromatography on sugar, and the order is chlorophyll *b* elutes first followed by chlorophyll *a*. To an Me<sub>2</sub>CO/H<sub>2</sub>O solution of chlorophylls 200ml of iso-octane are added, and the mixture shaken in a separating funnel and the H<sub>2</sub>O is carefully removed. The iso-octane layer is dried (Na<sub>2</sub>SO<sub>4</sub>) and applied onto a glass column (5cm diameter) dry packed with 1000ml of powdered sucrose which has been washed with 250ml of iso-octane. Elution with 0.5% of isopropanol in iso-octane gives chlorophyll *a*. Keeping the eluate overnight at 0° yields micro crystals which are collected by filtration or centrifugation (Yield 40mg). The UV<sub>EtOH</sub> has  $\lambda_{\text{max}}$  **660, 613, 577, 531, 498, 429** and 409 nm. *Note* that anhydrous chlorophylls can be easily enolised, epimerised or allomerised in dry polar organic solvents. [Anderson & Calvin *Nature* **194** 285 1962, Stoll & Weidemann *Helv Chim Acta* **16** 739 757 1933, NMR: Katz et al. *J Am Chem Soc* **90** 6841 1968, **85** 3809 1963 for *a* and *b*; ORD: Inhoffen et al. *Justus Liebigs Ann Chem* **704** 208 1967, Willstätter & Isler *Justus Liebigs Ann Chem* **390** 269, 233 1912, *Beilstein* **26** III/IV 3243.]

**Chlorophyll b** [519-62-0] **M 907.52, sinters at 86-92°, sinters at 170°, dec at 160-170°, m 183-185°, 190-195°,  $[\alpha]_D^{20} -267^\circ$  (Me<sub>2</sub>CO/MeOH),  $[\alpha]_{720}^{25} -133^\circ$  (MeOH/Pyridine 95:5).** See purification of chlorophyll *a*, and is separated from "*a*" by chromatography on sucrose [UV, IR: Stoll & Weidemann *Helv Chim Acta* **42** 679, 681 1959]. It forms red-black hexagonal bipyramids or four-sided plates from dilute EtOH and has been recrystallised from CHCl<sub>3</sub>/MeOH. It is soluble in MeOH, EtOH, EtOAc and insoluble in petroleum ether. [Dougherty et al. *J Am Chem Soc* **88** 5037 1966, *Beilstein* **26** III/IV 3787.]

**6-Chloropurine riboside (6-chloro-9-β-D-ribofuranosyl-9H-purine)** [2004-06-0] **M 286.7, m 158-162°(dec), 165-166°(sintering at 155°), 168-170°(dec),  $[\alpha]_D^{26} -45^\circ$  (c 0.8, H<sub>2</sub>O).** Purify the riboside by suspending the dry solid (~12 g) in hot MeOH (130 ml) and then adding enough hot H<sub>2</sub>O (~560ml) to cause solution, filter and set aside at 5° overnight. The colourless crystals of the riboside are filtered off, washed with Me<sub>2</sub>CO, Et<sub>2</sub>O and dried at 60°/0.1mm. More material can be obtained by evaporating the filtrate to dryness and recrystallisation of the residue from MeOH/H<sub>2</sub>O (2:1) (15ml/g). It has UV with  $\lambda_{\text{max}}$  at 264nm ( $\epsilon$  9140) in H<sub>2</sub>O. [Robins *Biochemical Preparations* **10** 145 1963, Baker et al. *J Org Chem* **22** 954 1957.]

**Chromomycin A<sub>3</sub>** [7059-24-7] **M 1183.3, m 185°dec,  $[\alpha]_D^{23} -57^\circ$  (c 1, EtOH).** Dissolve the antibiotic (10g) in EtOAc and add to a column of Silica Gel (Merck 0.05-0.2microns, 4x70cm) in EtOAc containing 1% oxalic acid. Elute with EtOAc+1% oxalic acid and check fractions by TLC. Pool fractions, wash with H<sub>2</sub>O thoroughly, dry and evaporate. Recrystallise the residue from EtOAc. The *heptaacetate* has **m** 214°,  $[\alpha]_D^{23} -20^\circ$  (c 1, EtOH). [Miyamoto et al. *Tetrahedron* **23** 421 1967, Harada et al. *J Am Chem Soc* **91** 5896 1969, *Beilstein* **17/5** V 673.]

**8*S*,9*R*-(-)-Cinchonidine** [485-71-2] **M 294.4, m 210.5°**,  $[\alpha]_{546}^{20} -127.5^{\circ}$  (c 0.5, EtOH),  $[\alpha]_{\text{D}}^{20} -109.2^{\circ}$  (EtOH), **pK<sub>1</sub><sup>15</sup> 4.17, pK<sub>2</sub><sup>15</sup> 8.4, pK<sub>1</sub><sup>20</sup> 5.80, pK<sub>2</sub><sup>20</sup> 10.03**. Crystallise cinchonidine from aqueous EtOH (prisms or plates). It is a strong base and readily forms a *mono- and di- hydrochloride* and other stable salts. It is slightly soluble in Et<sub>2</sub>O, but quite soluble in EtOH and CHCl<sub>3</sub>. For the *N-benzylcinchonidinium chloride* see [69257-04-1]. [Beilstein 23 III/IV 2824, 23/12 IV 406.]

**8*R*,9*S*-(+)-Cinchonine** [118-10-5] **M 294.4, m 265°**,  $[\alpha]_{546}^{20} +268^{\circ}$  (c 0.5, EtOH), **pK<sub>1</sub><sup>15</sup> 4.28, pK<sub>2</sub><sup>15</sup> 8.35, pK<sub>1</sub><sup>20</sup> 5.85, pK<sub>2</sub><sup>20</sup> 9.96**. Crystallise cinchonine from EtOH or diethyl ether (needles). Its solubility (w/v) is 1.7% in EtOH, 4% in boiling EtOH, 1% in CHCl<sub>3</sub> and is almost insoluble in H<sub>2</sub>O. It is a strong base and readily forms a *mono- and di- hydrochloride* and other stable salts. For the *N-benzylcinchoninium chloride* see [69221-14-3]. [Rabe *Justus Liebigs Ann Chem* 365 366, 371 1909, Beilstein 23 III/IV 2819, 2832.]

**Clonidine hydrochloride [Catapres, 2-(2,6-dichloroanilino)-2-imidazoline hydrochloride]** [4205-91-8] **M 266.6, m 305°, pK<sub>25</sub> 5.88 (free base)**. This antihypertensive is recrystallised from EtOH/Et<sub>2</sub>O and dried in a vacuum (solubility in H<sub>2</sub>O is 5%). The *free base* has **m 124-125°** and is recrystallised from hexane. [Jen et al. *J Med Chem* 18 90 1975, NMR: Jackman & Jen *J Am Chem Soc* 97 2811 1975.]

**Cloxacillin sodium salt (sodium 3-*o*-chlorophenyl-5-methyl-4-isoxazolyl penicillin monohydrate)** [642-78-4] **M 457.9, m 170°,  $[\alpha]_{\text{D}}^{20} +163^{\circ}$  (H<sub>2</sub>O pH 6.0-7.5), pK<sub>Est</sub> ~2.8 (COOH)**. Purify cloxacillin sodium salt by dissolving it in isoPrOH containing 20% of H<sub>2</sub>O, and diluting with isoPrOH to a water content of 5% and chilling. Recrystallise it again in this manner. The sodium salt is collected and dried at 40° in air to give the colourless *monohydrate*. It is soluble in H<sub>2</sub>O (5%), MeOH, EtOH, pyridine and ethylene glycol. [Doyle et al. *J Chem Soc* 5838 1963, Naylor et al. *Nature* 195 1264 1962.]

**(-)-Cocaine {ecogonine methyl ester benzoate, 2β-carbomethoxy-3-β-benzoxypyrane, methyl 1*R*-(*exo,exo*)-3-(benzoyloxy)-2-methyl-8-azabicyclo[3.2.1]octane-2-carboxylate}** [50-36-2] **M 303.4, m 98°, b 187-188°/0.1mm,  $[\alpha]_{\text{D}}^{20} -15.8^{\circ}$  (c 4, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{20} -35^{\circ}$  (50% EtOH), pK<sub>25</sub> 5.59 and 8.61 (8.39)**. (-)-Cocaine crystallises from EtOH and sublimes below 90° in a vacuum in an amorphous form. The *hydrochloride* crystallises from MeOH/Et<sub>2</sub>O with **m 195°** and  $[\alpha]_{\text{D}}^{20} -72^{\circ}$  (c 2 in H<sub>2</sub>O, pH 4.5), -78.5° (50% aqueous EtOH). [Sam & Reynolds *J Chem Soc* 97 1335 1910, Tufariello et al. *J Am Chem Soc* 101 2435 1979.] **α-Cocaine** is the (+) enantiomer. [Beilstein 22 I 547, 22 II 150.]

**Cocarboxylase tetrahydrate (aneurine pyrophosphoric acid tetrahydrate, thiamine pyrophosphoric acid tetrahydrate)** [136-09-4] **M 496.4, m 220-222°(sinters at 130-140°), 213-214°, pK<sub>Est(1)</sub> ~2, pK<sub>Est(2)</sub> ~6, pK<sub>Est(3)</sub> ~9**. Cocarboxylase tetrahydrate recrystallises from aqueous Me<sub>2</sub>CO. [Wenz et al. *Justus Liebigs Ann Chem* 618 210 1958, UV: Melnick *J Biol Chem* 131 615 1939, X-ray: Carlisle & Cook *Acta Cryst (B)* 25 1359 1969.] The *hydrochloride salt* has **m 242-244°(dec), 241-243°(dec) or 239-240°(dec)** and crystallises from aqueous HCl/EtOH, EtOH containing HCl or HCl/Me<sub>2</sub>CO. [Weijlard *J Am Chem Soc* 63 1160 1941, Synthesis: Weijlard & Tauber *J Am Chem Soc* 60 2263 1938, Beilstein 27 III/IV 1777.]

**Codeine** [76-57-3] **M 299.4, m 151-154°, 154-156°,  $[\alpha]_{\text{D}}^{20} -138^{\circ}$  (in EtOH), pK<sub>15</sub> 6.06(OH), pK<sub>25</sub> 8.21 (NMe)**. Codeine crystallises from water or aqueous EtOH. Dry it at 80°. Evaporation of a CHCl<sub>3</sub> extract gives a colourless glass which crystallises on scratching. It crystallises from H<sub>2</sub>O as the *monohydrate*, **m 157-158.5°**, and has  $[\alpha]_{\text{D}}^{25} -136^{\circ}$  (c 2.8, EtOH). The *hydrobromide* crystallises in needles from H<sub>2</sub>O, and effervesces at 151-160°, solidifies and remelts with extensive decomposition at 273-278°. It sublimes at 100°/0.03mm. [Gates *J Am Chem Soc* 75 4340 1953, Dauben et al. *J Org Chem* 44 1567 1979, Beilstein 27 II 136, 27 III/IV 2228.]

**Coenzyme A trihydrate** [85-61-0] **M 821.6, pK<sub>1</sub> 4.0 (adenine NH<sub>2</sub>), pK<sub>2</sub> 6.5 (PO<sub>4</sub>H), pK<sub>3</sub> 9.6 (SH)**. The white powder is best stored in an inert atmosphere in the dark in sealed ampoules after drying *in vacuo* over P<sub>2</sub>O<sub>5</sub> at 34°. It has UV with  $\lambda_{\text{max}}$  at 259 nm ( $\epsilon$  16,800) in H<sub>2</sub>O. [Buyske et al. *J Am Chem Soc* 76 3575 1954.] It is soluble in H<sub>2</sub>O but insoluble in EtOH, Et<sub>2</sub>O and Me<sub>2</sub>CO. It is readily oxidised in air and is best kept as the more stable *trilithium salt* [Moffat & Khorana *J Am Chem Soc* 83 663 1961; see also Beinert et al. *J Biol Chem* 200 384 1953, De Vries et al. *J Am Chem Soc* 72 4838 1950, Gregory et al. *J Am Chem Soc* 74 854 1952 and

Baddiley *Adv Enzymol* **16** 1 1955]. [*Beilstein* **26** III/IV 3663.]

**Coenzyme Q<sub>0</sub>** (2,3-dimethoxy-5-methyl-1,4-benzoquinone, 3,4-dimethoxy-2,5-tolu-quinone, fumigatin methyl ether) [605-94-7] **M 182.2, m 56-58°, 58-60°, 59°**. It crystallises in red needles from petroleum ether (b 40-60°) and sublimes at high vacuum at a bath temperature of 46-48° [Ashley et al. *J Chem Soc* 441 1938, UV in EtOH: Vischer *J Chem Soc* 815 1953, UV in cyclohexane: Morton et al. *Helv Chim Acta* **41** 2343 1858, Aghoramurthy et al. *Chem Ind (London)* 1327 1954]. [*Beilstein* **8** IV 2721.]

**Coenzyme Q<sub>4</sub>** (Ubiquinone-4, 2,3-dimethoxy-5-methyl-6-[3,7,11,15-tetramethyl-hexadeca-2*t*,6*t*,10*t*,14-tetraenyl]-1,4-benzoquinone) [4370-62-1] **M 454.7, m 30°, 33-45°**. It is a red oil which can be purified by chromatography on SiO<sub>2</sub> plates and eluted with Et<sub>2</sub>O/hexane. The purity is checked by HPLC (silica column using 7% Et<sub>2</sub>O/hexane). It has UV with  $\lambda_{\max}$  at 270 nm ( $\epsilon$  14,800) in petroleum ether. [NMR and MS: Naruta *J Org Chem* **45** 4097 1980, cf Morton *Biochemical Spectroscopy* (Adam Hilger, London, 1975) p 491]. It has also been dissolved in MeOH/EtOH (1:1 v/v) and kept at 5° until crystals appear [Lester & Crane *Biochim Biophys Acta* **32** 497 1958].

**Coenzyme Q<sub>9</sub>** (Ubiquinone-9, 2,3-dimethoxy-5-methyl-6-[3,7,11,15,19,23,27,31,35-nonamethyl-hexatriaconta-2*t*,6*t*,10*t*,14*t*,18*t*,22*t*,26*t*,30*t*,34-nonaenyl]-1,4-benzoquinone) [303-97-9] **M 795.3, m 40.5-42.5°, 44-45°, 45°**. The yellow crystals are purified by recrystallisation from petroleum ether and by chromatography on SiO<sub>2</sub> plates and eluted with Et<sub>2</sub>O/hexane. The purity can be checked by HPLC (silica column using 7% Et<sub>2</sub>O/hexane). It has UV with  $\lambda_{\max}$  at 270nm ( $\epsilon$  14,850) in petroleum ether. [NMR and MS: Naruta *J Org Chem* **45** 4097 1980, Le et al. *Biochem Biophys Acta* **32** 497 1958, cf Morton *Biochemical Spectroscopy* (Adam Hilger, London, 1975) p 491, IR: Lester et al. *Biochim Biophys Acta* **33** 169 1959, UV: Rüegg et al. *Helv Chim Acta* **42** 2616 1959, Shunk *J Am Chem Soc* **81** 5000 1959, *Beilstein* **8** IV 3313.]

**Coenzyme Q<sub>10</sub>** (Ubiquinone-10, 2,3-dimethoxy-5-methyl-6-[3,7,11,15,19,23,27,31,35,-39-decamethyl-tetraconta-2*t*,6*t*,10*t*,14*t*,18*t*,22*t*,26*t*,30*t*,34*t*,38-decaenyl]-1,4-benzoquinone) [303-98-0] **M 795.3, m 48-49°,49°, 49.5-50.5°, 50°**. Purify it by recrystallisation from EtOH, EtOH/Me<sub>2</sub>CO or Et<sub>2</sub>O/EtOH and by chromatography on silica gel using isoPrOH/Et<sub>2</sub>O (3:1) to give orange crystals. It has UV with  $\lambda_{\max}$  at 270nm ( $\epsilon$  15,170) in petroleum ether. [Terao et al. *J Org Chem* **44** 868 1979, NMR and MS: Naruta et al. *J Org Chem* **45** 4097 1980, IR: Lester et al. *Biochem Biophys Acta* **42** 1278 1959, NMR: Planta et al. *Helv Chim Acta* **42** 1278 1959; Morton *Biochemical Spectroscopy* (Adam Hilger, London, 1975) p 491].

**Colcemide (Demecocine)** [477-30-5] **M 371.4, m 182-185°, 183-185°, 186°,  $[\alpha]_{\text{D}}^{20}$  -129° (c 1, CHCl<sub>3</sub>)**. Colcemide is purified by chromatography on silica and eluting with CHCl<sub>3</sub>/MeOH (9:1), and by recrystallisation from EtOAc/Et<sub>2</sub>O to form yellow prisms. Its UV in EtOH has  $\lambda_{\max}$  at 243nm ( $\epsilon$  30,200) and 350nm ( $\epsilon$  16,300). [Synthesis, IR, NMR, MS: Capraro & Brossi *Helv Chim Acta* **62** 965 1979, *Beilstein* **8** IV 3319.]

**Colchicine** {*N*-[(7*S*)-1,2,3,10-tetramethoxy-5,6,7,9-tetrahydro-9-oxo-benzo[*a*]heptalen-7-yl]-acetamide} [64-86-8] **M 399.5, m 150-160° (dec), 155-157°(dec), 156-157°,  $[\alpha]_{\text{D}}^{20}$  -570° (c 1, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{20}$  -443° (c 1.7, H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{20}$  -120° (c 1, CHCl<sub>3</sub>), pK<sup>20</sup> 1.85 and 12.4**. Commercial material contains up to 4% desmethylcolchicine. Purify colchicine by chromatography on alumina and eluting with CHCl<sub>3</sub> [Ashley & Harris *J Chem Soc* 677 1944]. Alternatively, an acetone solution on alkali-free alumina has been used, and eluting with acetone [Nicholls & Tarbell *J Am Chem Soc* **75** 1104 1953]. It crystallises as yellow needles from EtOAc or CHCl<sub>3</sub> with solvent of crystallisation which can be removed at ~70°. It is soluble in Et<sub>2</sub>O (0.5%), \*C<sub>6</sub>H<sub>6</sub> (1%) and H<sub>2</sub>O (4%). It is sold as “Colgout” for the treatment of gout and binds to tubulin. [Schreiber et al. *Helv Chim Acta* **44** 540 1961, Scott et al. *Tetrahedron* **21** 3605 1965, van Tamelen et al. *Tetrahedron* **14** 8 1961, *Beilstein* **14** IV 946.]

**Compactin (Mevastatin)** [73573-88-3] **M 390.5, m 151-153°, 152°,  $[\alpha]_{\text{D}}^{22}$  +283° (c 0.48, acetone)**. Purify compactin by recrystallisation from aqueous EtOH. Its UV (EtOH) has  $\lambda_{\max}$  at 230, 237 and 246nm (log  $\epsilon$  4.28, 4.30 and 4.11); and its IR (KBr) has  $\nu_{\max}$  3520, 1750 (lactone CO) and 1710 (CO ester) cm<sup>-1</sup>. [Clive et al. *J Am Chem Soc* **110** 6914 1988, Review: Rosen & Heathcock *Tetrahedron* **42** 4909 1986, IR, NMR, MS: Brown et al.

*J Chem Soc Perkin Trans I* 1165 1976.] It is a potent inhibitor of 3-hydroxy-3-methylglutaryl-CoA reductase (HMG-CoA reductase), is an enzyme in cholesterol biosynthesis, and lowers cholesterol levels [Brown et al. *J Biol Chem* **253** 1121 1978, Nakamura & Ableles *Biochemistry* **24** 1364 1985, *Beilstein* **18/3** V 145].

**Crotaline** (monocrotaline, 12,13-dihydroxy-(13 $\beta$ -14 $\beta$ H)-14,19-dihydro-20-norcrotalanan-11,15-dione) [315-22-0] **M 325.4**, **m 196-197<sup>o</sup>(dec)**, **197-198<sup>o</sup>(dec)**, **203<sup>o</sup>(dec)**,  $[\alpha]_D^{20}$  -55<sup>o</sup> (c 1, EtOH). Crotaline forms prisms from absolute EtOH and recrystallises also from CHCl<sub>3</sub>. Its UV(96% EtOH) has  $\lambda_{\max}$  at 217nm (log  $\epsilon$  3.32). [Adams et al. *J Am Chem Soc* **74** 5612 1952, Culvenor & Smith *Aust J Chem* **10** 474 1957.] The hydrochloride has **m 212-214<sup>o</sup>** (from MeOH/Et<sub>2</sub>O) and  $[\alpha]_D^{25}$  -38.4<sup>o</sup> (c 5, H<sub>2</sub>O) [Adams & Gianturco *J Am Chem Soc* **78** 1922 1956]. The picrate has **m 230-231.5<sup>o</sup>(dec)** [Adams et al. *J Am Chem Soc* **74** 5614 1952]. [*Beilstein* **27** III/IV 6660.]

**Curcumin** [bis(4-hydroxy-3-methoxycinnamoyl)methane, 1,7-bis(4-hydroxy-3-methoxyphenyl)(1E,6E)-1,6-heptadiene-3,5-one] [458-37-7] **M 368.4**, **m 182<sup>o</sup>**, **183<sup>o</sup>**, **184-195<sup>o</sup>**, **184-186<sup>o</sup>**, **pK<sub>1</sub> 7.80**, **pK<sub>2</sub> 8.5** and **pK<sub>3</sub> 9.0 (H<sub>2</sub>O)**. Purify curcumin by recrystallising it from MeOH, EtOH or acetic acid (yellow, orange-yellow prisms) and drying it *in vacuo*. It is best separated from related compounds by HPLC (see refs below). It is sparingly soluble in H<sub>2</sub>O but soluble in AcOH and in alkaline buffer. In solution at pH 1-7 the colour is yellow and at pH 7-9 the colour is brownish-red or deep red. It was commercialised as a pH indicator under the name of "curcuma paper". Curcumin is a phenolic substance which occurs naturally in *Curcuma domestica*, *C. xanthorrhiza*, *C. aromatica*, *C. longa*, and is used in Indian curry cooking as *turmeric*. It has many physiological properties including anti-oxidant, anti-inflammatory, potent anti-tumour, and is an inhibitor of several enzymes (EGFR tyrosine kinase, I $\kappa$ B kinase, nitric oxide synthase, cyclooxygenase, lipoxygenase). It penetrates cell membranes accumulating in plasma membranes, nuclear envelope and endoplasmic reticulum. Its UV-VIS has  $\lambda_{\max}$  (log $\epsilon$ )(EtOH) at 268 (4.09) and 430nm (4.74), (dioxane) at 265 (4.18) and 420nm (4.37) and (40% aqueous THF) at 429nm (4.780), with fluorescence maxima at  $\lambda_{\text{excit}}$  433nm and  $\lambda_{\text{emis}}$  511nm in 40% aqueous THF. Its IR (KBr) has  $\nu_{\max}$  at 3400 (br), 1625, 1600, 1500, 1275, 1025, 960 cm<sup>-1</sup>. The <sup>1</sup>H NMR [400MHz, CDCl<sub>3</sub>] has  $\delta$  at 6.06 (1H, s, H-4), 16.41 (1H, br s, OH), 7.57 (2H, d, *J* = 16.0Hz, H-2,6), 6.75 (2H, d, *J* = 16.0Hz, H-1,7), 7.32 (2H, d, *J* = 2.0Hz, H-2',2''), 9.64 (2H, s, OH-4',4''), 6.85 (2H, d, *J* = 8.1Hz, H-5',5''), 7.16 (2H, dd, *J* = 2.0, 8.1Hz, H-6',6''), 3.85 (6H, s, 3',3'' OMe) (from TMS); and the <sup>13</sup>C NMR [100MHz, CDCl<sub>3</sub>] has  $\delta$  at 100 (C-4), 183.2 (C-3,5), 121.1 (C-2,6), 140.7 (C-1,7), 126.4 (C-1',1''), 111.5 (C-2',2''), 148.0 (C-3',3''), 149.4 (C-4',4''), 115.8 (C-5',5''), 123.0 (C-6',6''), 55.7 (OCH<sub>3</sub>) (from TMS). It complexes strongly with boron, iron, copper and nickel, but weakly with calcium and magnesium. As a 0.1% solution in AcOH it is used for the spectrophotometric determination of Boron as a complex with  $\lambda_{\max}$  550nm ( $\epsilon$ 180,000). [Dyrssen et al. *Analyt Chim Acta* **60** 139 1972, Spicer & Strickland *Analyt Chim Acta* **18** 231 1958, Roughley & Whiting, *J Chem Soc, Perkin Trans I* 2379 1973, Zsila et al. *Tetrahedron Asymm* **14** 2433 2003, cf Cooke & Segal *Aust J Chem* **8** 107 1955, *Beilstein* **8** H 554, **8** I 757, **8** II 588, **8** III 4312, **8** IV 3697.]

The *diacetyl* derivative has **m 171-172<sup>o</sup>** (from EtOH) and the *dibenzoyl* derivative has **m 210<sup>o</sup>** (from EtOH or C<sub>6</sub>H<sub>6</sub>\*).

The *dimethyl* derivative [1,7-bis(3,4-dimethoxyphenyl)(1E, 6E)-1,6-heptadiene-3,5-one] has **m 128-130<sup>o</sup>**, **M 396.2**; the UV has  $\lambda_{\max}$  at 427nm, and the <sup>1</sup>H NMR [200MHz, CDCl<sub>3</sub>] has  $\delta$  at 3.92 (6H, s, 2 x CH<sub>3</sub>O-Ph), 3.93 (6H, s, 2 x CH<sub>3</sub>O-Ph), 5.82 (1H, s, H-4), 6.45 (2H, d, *J* = 13Hz, H-1,2), 6.86 (6H, m, Arom-H), 7.57 (2H, d, *J* = 13Hz, H-6,7) [Nurfina et al. *Eur J Med Chem* **32** 321 1997].

**5'-Methoxycurcumin** [1-(4-hydroxy-3,5-dimethoxyphenyl)-7-(4-hydroxy-3-methoxyphenyl)-(1E,6E)-1,6-heptadiene-3,5-one] **M 398.4**, has **m 145-146<sup>o</sup>** (from MeOH), and was isolated as a yellow powder from *C. xanthorrhiza* with UV  $\lambda_{\max}$  at 429nm in MeOH.

**1-Hydroxycurcumin** [1-hydroxy-1,7-bis(4-hydroxy-3-methoxyphenyl)-(6E-6-heptene-3,5-dione)] **m 84-88<sup>o</sup>** was isolated from the same source, had UV  $\lambda_{\max}$  at 372nm in MeOH and had a weaker antioxidant activity than the other curcumins towards the autoxidation of linoleic acid in a water-alcohol system [Masuda et al. *Phytochemistry* **31** 3645 1992]. An optically active form, **(1 $\xi$ )-1-hydroxy-1,7-bis(4-hydroxy-3-methoxyphenyl)-(1E,6E)-1,6-heptadiene-3,5-one**, has **m 92.0-96.0<sup>o</sup>**, with  $[\alpha]_D^{20}$  +12.2<sup>o</sup> (c 0.06, EtOH, configuration unknown), **M 386.4**. Its UV has  $\lambda_{\max}$  nm ( $\epsilon$ )(MeOH) at 230 (sh), 260 (21000), 283 (18000).

370 (56000); and its IR (CDCl<sub>3</sub>) has  $\nu_{\max}$  at 965, 1000, 1120, 1260, 1325, 1375, 1460, 1560, 1600, 1620, 1650, 2850, 2900, 3525 cm<sup>-1</sup>. Its <sup>1</sup>H NMR [270MHz, CDCl<sub>3</sub>] has  $\delta$  at 2.65(1H, dd,  $J = 17, 3.5$ Hz), 2.93 (1H, dd,  $J = 17, 14$ Hz), 3.93 (3H, s), 3.95 (3H, s), 5.39 (1H, dd,  $J = 14, 3.5$ Hz), 5.60 (1H, s), 5.80 (1H, br s, disappeared on addition of D<sub>2</sub>O), 5.86 (1H, br s, disappeared on addition of D<sub>2</sub>O), 6.44 (1H, d,  $J = 16$ Hz), 6.90 (1H, d,  $J = 8$ Hz), 7.00 (4H, br s), 7.03 (1H, dd,  $J = 8, 2$ Hz), 7.30 (1H, d,  $J = 16$ Hz) and  $R_F$  (Silica gel 60 F<sub>254</sub>, C<sub>6</sub>H<sub>6</sub>\*/EtOAc 1:1) of 0.2. When this *hydroxyheptadiene* was heated at 130°/15mmHg/10 hours it gave *curcumin* in 79% yield. (**3S,5S**) - **1, 7 - Bis (4 - hydroxy-3-methoxyphenyl) heptane-3, 5-diol** [colourless needles **m 132-134**°, with  $[\alpha]_D^{20} -18.5$ ° (c 0.26, EtOH),  $R_F$  0.1] was also isolated from the rhizomes of *C. xanthorrhiza* (Zingiberaceae) and characterised; and (**3R,5R**)-**1-(4-hydroxyphenyl)-7-phenylheptane-3,5-diol** [colourless needles **m 109-111**°, with  $[\alpha]_D^{20} +8.3$ ° (c 0.06, CDCl<sub>3</sub>),  $R_F$  0.3] was isolated from the rhizomes of *Alpinia officinarum* (Zingiberaceae) and characterised; the absolute configurations were determined from the circular dichroism (CD) spectra of the 3,5-bis-*p*-dimethylaminobenzoate esters of their methylated derivatives, and applying the *exciton chirality rule* [Uehara et al. *Chem Pharm Bull Jpn* **38** 3298 1987].

**Desmethoxycurcumin** [22608-11-3] **M 338.3, m 173-175**° (also reported are **168**° and **172**°) is a contaminant of curcumin, isolated from the same source as curcumin, best purified by HPLC and crystallises from EtOH as an orange-yellow powder. Its UV-VIS has  $\lambda_{\max}$  (log  $\epsilon$ ) (MeOH) at 419nm (4.706) and (40% aqueous THF) 425nm (4.806) with fluorescence maxima at  $\lambda_{\text{excit}}$  428nm and  $\lambda_{\text{emis}}$  505nm in 40% aqueous THF. The <sup>1</sup>H NMR [400MHz, CDCl<sub>3</sub>] has  $\delta$  at 5.97 (1H, s, H-4), 7.60 (2H, d,  $J = 16.0$ Hz, H-2,6), 6.69 (1H, d,  $J = 16.0$ Hz, H-1 or 7), 6.64 (1H, d,  $J = 16.0$ Hz, H-7 or 1), 7.34 (1H, d,  $J = 1.7$ Hz, H-2' or 2''), 6.90 (1H, d,  $J = 8.0$ Hz, H-2'' or 2'), 7.56 (1H, d,  $J = 8.0$ Hz, H-5'), 6.88 (1H, d,  $J = 8.0$ Hz, H-5''), 7.27 (1H, dd,  $J = 8, 1.7$ Hz, H-6'), 6.9 (1H, d,  $J = 8$ Hz, H-6''), 3.92 (3H, s, OCH<sub>3</sub>) (from TMS); and the <sup>13</sup>C NMR [100MHz, CDCl<sub>3</sub>] has  $\delta$  at 101.6 (C-4), 184.4 and 184.5 (C-3,5), 121.1 and 122.3 (C-2,6), 141.4 and 141.0 (C-1,7), 128.2 and 127.7 (C-1',1''), 111.5 and 130.9 (C-2',2''), 148.8 and 116.8 (C-3',3''), 150.0 and 160.5 (C-4',4''), 116.2 and 116.8 (C-5',5''), 123.8 and 130.9 (C-6',6''), 56.3 (OCH<sub>3</sub>) (from TMS).

Similarly **bis-desmethoxy-curcumin** [22608-12-4] **M 308.3, m 224**° (also reported are **216-218**° and **222**°) is a contaminant of curcumin, occurring in the same source, best purified by HPLC and crystallises as the *hydrate* in yellow plates from EtOH. Its UV has  $\lambda_{\max}$  (log  $\epsilon$ )(MeOH) at 414nm (4.675) and (40% aqueous THF) 420nm (4.702) with fluorescence maxima at  $\lambda_{\text{excit}}$  425nm and  $\lambda_{\text{emis}}$  501nm in 40% aqueous THF. The <sup>1</sup>H NMR [400MHz, CDCl<sub>3</sub>] has  $\delta$  at 6.03 (1H, s, H-4), 16.4 (1H, br s, 3-OH), 7.56 (2H, d,  $J = 15.9$ Hz, H-2,6), 7.56 (2H, d,  $J = 15.9$ Hz, H-1,7), 6.84 (2H, d,  $J = 8.2$ Hz, H-2',2''), 7.56 (2H, d,  $J = 8.2$ Hz, H-3',3''), 10.03 (2H, s, 4',4''-OH), 7.56 (2H, d,  $J = 8.2$ Hz, H-5',5''), 8.64 (2H, d,  $J = 8.2$ Hz, H-6',6'') (from TMS); and the <sup>13</sup>C NMR [100MHz, CDCl<sub>3</sub>] has  $\delta$  at 100.9 (C-4), 183.2 (C-3,5), 120.8 (C-2,6), 140.3 (C-1,7), 125.8 (C-1',1''), 130.3 (C-2',2''), 115.9 (C-3',3''), 159.8 (C-4',4''), 115.9 (C-5',5''), 130.3 (C-6',6'') (from TMS). [Jayaprakasha et al. *J Agric Food Chem* **50** 3668 2002, Inoue et al. *J Agric Food Chem* **56** 9328 2008, (HPLC, UV, fluorescence) Rouseff *J Food Sci* **53** 1823 1988, (MS) Jiang et al. *Rapid Commun Mass Spectrom* **20** 1001 2006, (HPLC, MS) Jiang et al. *J Chromatogr A* **1111** 21 2006, (HPLC, MS) Hiserodt et al. *J Chromatogr A* **740** 51 1996, (spectroscopy) Tonnesen et al. *Pharmazie* **50** H.10 1995, (IR) Tanaka et al. *J Agric Food Chem* **56** 9787 2008, (spectroscopy) Péret-Almeida et al. *Food Research International* **38** 1039 2005. (cytotoxicity) Ishida et al. *Bioorg Med Chem* **10** 3481 2002 (anti-inflammatory, antioxidant and anti-amyloid activity) Yang et al. *J Biol Chem* **280** 5892 2005.]

The keto-enol tautomerism of curcumin derivatives was shown to be a property that can be employed for amyloid detection in Alzheimer's disease, and may be exploited for the design of amyloid-binding agents for the therapy of the disease [Yanagisawa et al. *Biomaterials* **31** 4179 2010]. Over the years, curcumin derivatives have been screened for a variety of diseases (see above) including anti-HIV, and although few have displayed very potent anticancer activity, they are still being considered for development as future anticancer agents. A review on their potential as anticancer agents has been published recently. [Agrawal & Mishra *Med Res Rev* **30** 818 2010.] The above information under *Curcumin* was kindly supplied by Prof. Hiroyasu Taguchi, Shiga University of Medical Science, Molecular Neuroscience Research Centre, Japan.

**Cytidine** [65-46-3] **M 243.2, m 210-220**°(dec), **230**° (dec), **251-252**° (dec),  $[\alpha]_D^{20} +37$ ° (c 9, H<sub>2</sub>O),  $[\alpha]_D^{20} +29$ ° (c 9, H<sub>2</sub>O), **pK<sub>25</sub> 3.85**. Cytidine crystallises from 90% aqueous EtOH. It has also been converted to *sulfate* by dissolving (~200mg) in a solution of EtOH (10ml) containing H<sub>2</sub>SO<sub>4</sub> (50mg), whereby the salt crystallises out.

It is collected, washed with EtOH and dried for 5 hours at 120°/0.1mm. The *sulfate* has **m** 225°. The *free base* is obtained by shaking the salt solution with a weak ion-exchange resin, filtering, evaporating and recrystallising the residue from EtOH as before. [Fox & Goodman *J Am Chem Soc* **73** 3256 1956, Fox & Shugar *Biochim Biophys Acta* **9** 369 1952; see Prytsas & Sorm in *Synthetic Procedures in Nucleic Acid Chemistry* (Zorbach & Tipson Eds) **Vol 1** 404 1973.] [*Beilstein* **25** III/IV 3667.]

**Cytisine**, See entry in “Heterocyclic Compounds”, Chapter 4.

**Cytochalasin B (from dehydrated mould matter)** [14930-96-2] **M 479.6, m 218-221°**. Purify it by MeOH extraction, reverse phase C18 silica gel batch extraction by selective elution with 1:1 v/v hexane/tetrahydrofuran, crystallisation, subjected to TLC and recrystallisation [Lipski et al. *Anal Biochem* **161** 332 1987]. It is soluble in EtOH (3.6%), Me<sub>2</sub>CO (1%), Me<sub>2</sub>SO (37%) and Me<sub>2</sub>NCHO (49%) at 24°, and can be crystallised from the first two solvents. It interferes with cellular movement [Korm *Physiol Reviews* **62** 672 1982].

**Cytosine (4-amino-2-hydroxypyrimidine)** [71-30-7] **M 111.1, m 313° (hydrate, dec), 320-325° (anhydrous, dec), pK<sub>1</sub><sup>25</sup> 4.6, pK<sub>2</sub><sup>25</sup> 12.1**. Cytosine crystallises from H<sub>2</sub>O as the *monohydrate* which loses water on heating above 100°. Its solubility in H<sub>2</sub>O is 0.77%. Its UV has λ<sub>max</sub> at 267nm (ε 6,100) in H<sub>2</sub>O pH 8.8 and 275nm (ε 10,400) in 0.1N HCl. [Hilbert & Johnson *J Am Chem Soc* **52** 1152 1930, Hilbert et al. *J Am Chem Soc* **57** 552 1935, *Beilstein* **25** III/IV 3654.]

**Cytosine-1-β-O-arabinofuranoside (Cytarabin)** [147-94-4] **M 243.2, m ~220°(dec), 212-213.5°, [α]<sub>D</sub><sup>20</sup> +155° (c 1, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 4.3**. Purify cytarabin by recrystallisation from aqueous EtOH or a large volume of H<sub>2</sub>O (it solubility at ~20° is 5%). It has λ<sub>max</sub> at 212 and 279nm at pH 2 and 272nm at pH 12. It is an acute leukaemic agent. [Walwick et al. *Proc Chem Soc (London)* 84 1959, *Beilstein* **25** III/IV 3669.]

**Demeclocycline hydrochloride (7-chloro-6-demethyltetracycline hydrochloride, Clortetrin)** [64-73-3] **M 501.3, m 174-178°(dec, for sesquihydrate), [α]<sub>D</sub><sup>25</sup> -258° (c 0.5, 0.1N H<sub>2</sub>SO<sub>4</sub>), pK<sub>1</sub><sup>25</sup> 4.45 [H<sub>2</sub>O-Me<sub>2</sub>NCHO (1:1)]**. Crystallise the salt from EtOH/Et<sub>2</sub>O or H<sub>2</sub>O and dry it in air [McCormick et al. *J Am Chem Soc* **79** 4561 1957, Dobrynin et al. *Tetrahedron Lett* 901 1962]. [*Beilstein* **14** IV 2625.]

**2'-Deoxyadenosine (adenine 2'-deoxyriboside)** [16373-93-6] **M 269.3, m 187-189°, 189-191°, [α]<sub>D</sub><sup>20</sup> -25° (c 0.5, H<sub>2</sub>O), [α]<sub>589</sub><sup>25</sup> -26°, [α]<sub>310</sub><sup>25</sup> -206° (c 0.5, H<sub>2</sub>O), pK<sub>20</sub><sup>20</sup> 3.79**. Purify it by recrystallisation from H<sub>2</sub>O (hydrated crystals; solubility of the *monohydrate* is 1.1% in H<sub>2</sub>O at 20°). It has UV with λ<sub>max</sub> at 258nm (pH 1), 260nm (pH 7) and 261nm (pH 13). [Ness & Fletcher *J Am Chem Soc* **81** 4752 1959, Walker & Butler *Can J Chem* **34** 1168 1956.] The 3',5'-*O*-diacetyl derivative has **m** 151-152° (from EtOAc/petroleum ether). [*Beilstein* **26** III/IV 3589.]

**3'-Deoxyadenosine (Cordycepin, adenine 3'-deoxyriboside)** [73-03-0] **M 251.2, m 225-226°, 225-229°, [α]<sub>D</sub><sup>20</sup> -47° (H<sub>2</sub>O), pK<sub>Est</sub> ~4.8**. 3'-Deoxyadenosine forms needles from EtOH, *n*-BuOH and *n*-PrOH, and a *monohydrate* from H<sub>2</sub>O. It has UV with λ<sub>max</sub> at 260nm (ε 14,600) in EtOH. The *picrate* has **m** 195°(dec, yellow crystals from H<sub>2</sub>O). [Kaczka et al. *Biochim Biophys Acta* **14** 456 1964, Todd & Ulbricht *J Chem Soc* 3275 1960, Lee et al. *J Am Chem Soc* **83** 1906 1961, Walton et al. *J Am Chem Soc* **86** 2952 1964, *Beilstein* **26** III/IV 3594.]

**2'-Deoxycytidine monohydrate** [951-77-9, 652157-52-3 (*hydrate*), 207121-53-7] **M 245.2, m 119-200°, 207-209°, 213-215°, [α]<sub>D</sub><sup>25</sup> +78° (c 0.4, N NaOH), [α]<sub>D</sub><sup>23</sup> +57.6° (c 2, H<sub>2</sub>O), pK<sub>1</sub><sup>25</sup> 4.25**. Purify 2'-deoxycytidine by recrystallisation from MeOH/Et<sub>2</sub>O or EtOH and dry it in air. [NMR: Miles *J Am Chem Soc* **85** 1007 1963, UV: Fox & Shugar *Biochim Biophys Acta* **9** 369 1952.] The *hydrochloride* crystallises from H<sub>2</sub>O/EtOH and has **m** 174°(dec, 169-173°). [Walker & Butler *Can J Chem* **34** 1168 1956.] The *picrate* has **m** 208°(dec). [Fox et al. *J Am Chem Soc* **83** 4066 1961, *Beilstein* **25** III/IV 3662.]

**2'-Deoxycytidine 5'-monophosphoric acid (deoxycytidylic acid)** [1032-65-1] **M 307.2, m 170-172°(dec),**

183-184°(dec), 183-187°(dec),  $[\alpha]_D^{21} +35^\circ$  (c 0.2, H<sub>2</sub>O), pK<sub>1</sub> 4.6, pK<sub>2</sub> 6.6. Recrystallise the acid from H<sub>2</sub>O or aqueous EtOH and dry it in a vacuum. [Volkin et al. *J Am Chem Soc* **73** 1533 1951, UV: Fox et al. *J Am Chem Soc* **75** 4315 1953, IR: Michelson & Todd *J Chem Soc* 3438 1954, *Beilstein* **25** IV 3664.]

**2'-Deoxyguanosine monohydrate** (9-[2-deoxy-β-D-ribofuranosyl]guanidine) [961-07-9] M 285.3, m ca 200°(dec),  $[\alpha]_D^{20} +37.5^\circ$  (c 2, H<sub>2</sub>O),  $[\alpha]_D^{14} -47.7^\circ$  (c 0.9, N NaOH), pK<sub>Est(1)</sub>~ 3.3, pK<sub>Est(2)</sub>~ 9.2. 2'-Deoxyguanosine recrystallises from H<sub>2</sub>O as the monohydrate. [Brown & Lythgoe *J Chem Soc* 1990 1950, Levene & London *J Biol Chem* **81** 711 1929, **83** 793 1929, UV: Hotchkiss *J Biol Chem* **175** 315 1948, ORD: Levendahl & James *Biochim Biophys Acta* **26** 89 1957.] The 3',5'-di-O-acetyl derivative crystallises from aqueous EtOH with m 222°(dec), and  $[\alpha]_D^{18} -38^\circ$  (c 0.3, 10% aqueous EtOH) [Hayes et al. *J Chem Soc* 808, 813 1955]. [*Beilstein* **26** III/IV 3897.]

**2'-Deoxyinosine** [890-38-0] M 252.2, m 206°(dec), 218-220°(dec),  $[\alpha]_D^{25} -21^\circ$  (c 2, N NaOH),  $[\alpha]_D^{21.5} -21^\circ$  (c 1, H<sub>2</sub>O), pK<sub>Est(1)</sub>~8.9, pK<sub>Est(2)</sub>~12.4. Purify 2'-deoxyinosine by recrystallisation from H<sub>2</sub>O. [Brown & Lythgoe *J Chem Soc* 1990 1950, UV: MacNutt *Biochem J* **50** 384 1952, *Beilstein* **26** III/IV 2086.]

**5-Deoxy-5-(methylthio)adenosine** [2457-80-9] M 297.3, m 210-213°(dec), 211°, 212°, 213-214°.  $[\alpha]_D^{20} -23.7^\circ$  (c 0.02, pyridine),  $[\alpha]_D^{20} -8^\circ$  (c 1, 5% aqueous NaOH),  $[\alpha]_D^{25} +15^\circ$  (c 0.4-1.0, 0.3N aqueous AcOH), pK<sub>Est</sub>~3.5. Recrystallise it from H<sub>2</sub>O and sublime it at 200°/0.004mm. [v.Euler & Myrbäck *Hoppe Seyler's Z physiol Chem* **177** 237 1928, Weygand & Trauth *Chem Ber* **84** 633 1951, Baddiley et al. *J Chem Soc* 2662 1953.] The hydrochloride has m 161-162° [Kuhn & Henkel *Hoppe Seyler's Z Physiol Chem* **269** 41 1941]. The picrate has m 183°(dec) (from H<sub>2</sub>O). [*Beilstein* **26** III/IV 3675.]

**Deoxyribonucleic acid (from plasmids).** Purify plasmid DNA by two buoyant density ultracentrifugations using ethidium bromide-CsCl. The ethidium bromide is extracted with Et<sub>2</sub>O, and the DNA is dialysed against buffered EDTA and lyophilised. [Marmur & Doty *J Mol Biol* **5** 109 1962, Guerry et al. *J Bacteriol* **116** 1064 1973.] See "Introduction" in this Chapter.

**3'-Deoxythymidine** {2',3'-dideoxythymidine, 1-[(2*r*)-5*c*-hydroxymethyltetrahydro(2*r*)-furyl]-5-methylpyrimidine-2,4-dione} [3416-05-5] M 226.2, m 145°, 149-151°,  $[\alpha]_D^{26} +18^\circ$  (c 1, H<sub>2</sub>O), pK<sub>Est</sub>~ 9.2. Crystallise it from Me<sub>2</sub>CO/MeOH. [Michelson & Todd *J Chem Soc* 816 1955, *Beilstein* **24** III/IV 1297.]

**2'-Deoxyuridine** [1-(β-D-erythro-2-deoxypentofuranosyl)-1*H*-pyrimidine-2,4-dione] [951-78-0] M 228.2, m 163°, 163-163.5°, 165-167° 167°,  $[\alpha]_D^{26} +30^\circ$  (c 2, H<sub>2</sub>O),  $[\alpha]_D^{22} +50^\circ$  (c 1, N NaOH), pK<sub>25</sub> 9.3. 2'-Deoxyuridine forms needles from absolute EtOH or 95% EtOH. [Dekker & Todd *Nature* **166** 557 1950, Brown et al. *J Chem Soc* 3035 1958, NMR Jardetzky *J Am Chem Soc* **83** 2919 1961, Fox & Shugar *Biochim Biophys Acta* **9** 369 1952, UV: MacNutt *Biochem J* **50** 384 1952, *Beilstein* **24** III/IV 1200.]

**3'-Deoxyuridine** {1-[(2*R*)-5*c*-hydroxymethyltetrahydro(2*r*)furyl]-5-methylpyrimidin-2,4-dione, 2',3'-dideoxythymidine} [7057-27-4, 3416-05-5] M 226.2, m 149-151°,  $[\alpha]_D^{20} +18^\circ$  (c 1, H<sub>2</sub>O), pK<sub>Est</sub>~ 9.3. 3'-Deoxyuridine is recrystallised from Me<sub>2</sub>CO/MeOH and is dried in a vacuum. [Michelson & Todd *J Chem Soc* 816 1955.]

**Desthiobiotin (4*R*-cis-5-methyl-2-oxo-4-imidazolidinehexanoic acid)** [533-48-2] M 214.3, m 156-158°,  $[\alpha]_D^{20} +10.5^\circ$  (c 2, H<sub>2</sub>O), pK<sub>Est</sub>~2.8. Dissolve desthiobiotin in 0.5% Na<sub>2</sub>CO<sub>3</sub>, filter, acidify with HCl to Congo Red, concentrate to a small volume (2-3 ml) to give fine needles, filter it off and recrystallise it twice from H<sub>2</sub>O, m 157-158°. It also crystallises from 95% EtOH. The methyl ester crystallises from MeOH and sublimes at 100°/high vacuum, m 69-70°,  $[\alpha]_D^{27.5} +2.6^\circ$  (c 2, CHCl<sub>3</sub>). [Melville et al. *Science* **98** 497 1943, *J Am Chem Soc* **66** 1422 1944, *Beilstein* **25** III/IV 1543.]

**Di- and tri-carboxylic acids.** These are separated by anion-exchange chromatography. [Bengtsson & Samuelson *Anal Chim Acta* **44** 217b 1969.]

**7,8-Dihydrofolic acid (7,8-dihydropteroyl-L-glutamic acid, DHFA)** [4033-27-6] M 443.4, pK<sub>1</sub> 2.0 (basic

**10-NH), pK<sub>2</sub> 2.89 (2-NH<sub>2</sub>), pK<sub>3</sub> 3.45 (α-CO<sub>2</sub>H), pK<sub>4</sub> 4.0 (basic 5N), pK<sub>5</sub> 4.8 (γ-CO<sub>2</sub>H), pK<sub>6</sub> 9.54 (acidic 3NH).** DHFA is best purified by suspending (1g mostly dissolved) in ice-cold sodium ascorbate (300ml of 10% at pH 6.0, prepared by adjusting the pH of 30g of sodium ascorbate in 150ml of H<sub>2</sub>O by adding 1N NaOH dropwise using a glass electrode till the pH is 6.0). This gives a clear solution with pH ~5. While stirring at 0°, add N HCl dropwise slowly (0.1ml/min) until the pH drops to 2.8 when white birefringent crystals separate. These are collected by centrifugation (1000xg for 5 minutes), washed 3x with 0.001N HCl also by centrifugation and decantation. The residue is then dried in a vacuum (0.02mm) over P<sub>2</sub>O<sub>5</sub> (change the P<sub>2</sub>O<sub>5</sub> frequently at first) and KOH at 25° in the dark. After 24 hours the solid reaches constant weight.

For the assay of *dihydrofolate reductases* (see [9002-03-3] p 803): suspend ~66.5mg of DHFA in 10ml of 0.001M HCl containing 10mM dithiothreitol (DTT stock made from 154mg in 10ml H<sub>2</sub>O making 0.1M), shake well and freeze in 400μl aliquots. Before use, mix 400μl of this suspension with 0.1M DTT (200μl, also made in frozen aliquots), and the mixture is diluted with 200μl of 1.5M Tris-HCl pH 7.0 and 1.2ml of H<sub>2</sub>O (making a total volume of ~2ml) to give a clear solution. To estimate the concentration of DHFA in this solution, dilute 20μl of this solution to 1ml with 0.1M Tris-HCl pH 7.0 and read the OD at 282nm in a 1cm path length cuvette. ε at 282nm is 28,000M<sup>-1</sup>cm<sup>-1</sup>. [Futterman *Methods Enzymol* **6** 801 1963, Reyes & Rathod *Methods Enzymol* **122** 360 1986, *Beilstein* **26** III/IV 3934.]

**DL-erythro-Dihydrosphingosine (dl-erythro-2-amino-octadecan-1,3-diol) [3102-56-5] M 301.5, m 85-86°, 85-87°, pK<sub>Est</sub> ~ 8.8.** Purify it by recrystallisation from petroleum ether/EtOAc or CHCl<sub>3</sub>. The (±)-*N*-dichloroacetyl derivative has **m** 142-144° (from MeOH). [Shapiro et al. *J Am Chem Soc* **80** 2170 1958, Shapiro & Sheradsky *J Org Chem* **28** 2157 1963.] The D-isomer crystallises from petroleum ether/Et<sub>2</sub>O and has **m** 78.5-79°, [α]<sub>546</sub><sup>28</sup> +6° (CHCl<sub>3</sub>/MeOH, 10:1). [Grob & Jenny *Helv Chim Acta* **35** 2106 1953, Jenny & Grob *Helv Chim Acta* **36** 1454 1953, *Beilstein* **4** I 448, **4** II 757, **4** III 854, **4** IV 1887.]

**Dihydrostreptomycin sesquisulfate [5490-27-7] M 461.4, m 250°(dec), 255-265°(dec), [α]<sub>D</sub><sup>20</sup> -92.4° (c 1, H<sub>2</sub>O), pK<sub>Est(1)</sub> ~ 9.5 (NMe), pK<sub>Est(2,3)</sub> ~ 13.4 (guanidino).** It crystallises from H<sub>2</sub>O, MeOH, *n*-BuOH or methyl ethyl ketone. The crystals are not hygroscopic like the amorphous powder; however, both forms are soluble in H<sub>2</sub>O but the amorphous solid is about 10 times more soluble than the crystals. The *free base* also crystallises from H<sub>2</sub>O/Me<sub>2</sub>CO and has [α]<sub>D</sub><sup>26</sup> -92° (aqueous solution pH 7.0). [Solomons & Regina *Science* **109** 515 1949, Wolf et al. *Science* **109** 515 1949, McGilveray & Rinehart *J Am Chem Soc* **87** 4003 1956]. [*Beilstein* **18** III/IV 7538.]

**2,4-Dihydroxyimidazole (hydantoin, imidazolin-2,4-dione) [461-72-3] M 100.1, m 215°, 218-220°, 220°, pK<sup>25</sup> 9.15, pK<sup>18</sup> 9.16.** Well over the past hundred years many syntheses of hydantoin have been reported involving the reaction of alloxanic acid with HI [Baeyer *Justus Liebigs Ann Chem* **119** 127 1861 and **130** 159 1864], the cyclisation of urea derivatives [Harries & Weiss *Chem Ber* **33** 3418 1900 and *Justus Liebigs Ann Chem* **327** 366 1903], the cyclisation of ethyl *N*-carbamoylglycinate, accomplished with NaOMe in MeOH in 85% yield [Kavalek et al. *Col Czech Chem Commun* **51** 375 1986], and more recently by the acid catalysed condensation of glycine with KNCO [El-Deeb et al. *Eur J Med Chem* **45** 2516 2010]. In a later patent, glycine (75g, 1 mole) and urea (140g, 2.3 moles) were dissolved in H<sub>2</sub>O (120ml) and stirred under reflux for 12 hours. The mixture was acidified by careful addition of concentrated H<sub>2</sub>SO<sub>4</sub> while being cooled in an ice bath, then boiled for 1 hour, cooled to 0-5°, and the crystalline solid was filtered off, washed with cold H<sub>2</sub>O and dried. A further crop was obtained by evaporating the mother liquor. The combined solids were recrystallised from glacial AcOH (10 parts) to give pure *hydantoin* (71g, 70%) m 220-220°. [Alkaloida US patent 4647694 1987.] It can be purified by dissolving it in at least 2.5 parts of hot H<sub>2</sub>O (charcoal), filtering, acidifying to pH ~6, filtering again, collecting the solid, washing it with ice cold H<sub>2</sub>O, drying it, and recrystallising it from EtOH (needles) or from 10 parts of AcOH. Its solubility is 39.7g/L in H<sub>2</sub>O and 3.24g/L in EtOH at 25°. It has λ<sub>max</sub> (ε) at 215(400) and 245(60)nm in 0.01N EtOH/NaOH; 215(360) and 245(50)nm in EtOH/H<sup>+</sup>; and has ν<sub>max</sub> at 1697 (2CO) and 1776 (4CO) cm<sup>-1</sup> in a paraffin mull [West *J Biol Chem* **34** 187 1918, Crombie & Hooper *J Chem Soc* 3010 1955, *Beilstein* **24** H 242, **24** I 287, **24** II 127, **24** III/IV 1034].

**1,2-Dilauroyl-*sn*-glycero-3-phosphoethanolamine (±-dilauroyl-α-cephalin, 3-*sn*-phosphat-idylethanolamine 1,2-didodecanoyl) [59752-57-7] M 579.8, m 210°, pK<sub>Est(1)</sub> ~ 5.8 (PO<sub>4</sub>H), pK<sub>Est(2)</sub> ~ 10.5 (NH<sub>2</sub>).**

Recrystallise it from EtOH or tetrahydrofuran. [Bevan & Malkin *J Chem Soc* 2667 1951, IR: Bellamy & Beecher *J Chem Soc* 728 1953, *Beilstein* 4 IV 1417.]

**1,2-Dimyristoyl-sn-glycero-3-phosphocholine monohydrate (dimyristoyl-L- $\alpha$ -lecithin)** [18194-24-6 (1H<sub>2</sub>O)] **M 696.0**,  $[\alpha]_D^{24} +7.0$  (c 8, EtOH-CHCl<sub>3</sub> 1:1 for  $\alpha_1$  form), **pK<sub>Est</sub> ~ 5.8 (PO<sub>4</sub>)**. It has three forms  $\alpha_1$ ,  $\alpha_2$  and  $\beta'$ . Recrystallise it from aqueous EtOH or EtOH/Et<sub>2</sub>O. Its solubility at 22-23° in Et<sub>2</sub>O is 0.03%, in Me<sub>2</sub>CO it is 0.06% and in pyridine it is 1.3%. [Baer & Kates *J Am Chem Soc* 72 942 1950, Baer & Maurakas *J Am Chem Soc* 74 158 1952, IR: Marinetti & Stotz *J Am Chem Soc* 76 1347 1954.] The *S*-isomer with 1H<sub>2</sub>O is recrystallised from 2,6-dimethylheptan-4-one and has **m** 226-227° (sintering at 90-95°), and  $[\alpha]_D^{20} -7.0$  (c 6, MeOH/CHCl<sub>3</sub> 1:1). [Baer & Martin *J Biol Chem* 193 835 1951, *Beilstein* 4 IV 1463.]

**(±)-1,2-Dimyristoyl-sn-glycero-3-phosphoethanolamine (dimyristoyl- $\alpha$ -kephalin)** [998-07-2] **M 635.9**, **m 207°**, **pK<sub>Est(1)</sub> ~ 5.8 (PO<sub>4</sub>H)**, **pK<sub>Est(2)</sub> ~ 10.5 (NH<sub>2</sub>)**. Recrystallise the kephalin from EtOH [Bevan & Malkin *J Chem Soc* 2667 1951]. The *R*-isomer has **m** 195-196° (sintering at 130-135°) after recrystallisation from CHCl<sub>3</sub>/MeOH, and  $[\alpha]_D^{26} +6.7$  (c 8.5, CHCl<sub>3</sub>/AcOH 9:1). [Baer *Can J Biochem Physiol* 35 239 1957, Baer et al. *J Am Chem Soc* 74 152 1952, *Beilstein* 4 IV 1463.]

***S*(-)-1,2-Dipalmitin (*S*-1,2-dipalmitoyl-sn-glycerol)** [30334-71-5] **M 568.9**, **m 68-69°**  $[\alpha]_D^{20} -2.9$  (c 8, CHCl<sub>3</sub>), the ***R*(+)-isomer (2,3-dipalmitoyl-sn-glycerol)** [6076-30-8] has **m 67.5-68.2°**,  $[\alpha]_D^{20} -2.8$  (c 8, CHCl<sub>3</sub>). Crystallise *S*(-)-1,2-dipalmitin from chloroform/petroleum ether (b 40-60°) ~1:1.5. [*S*(-)-isomer: Baer & Kates *J Am Chem Soc* 72 942 1950, Hanahan & Vercamer *J Am Chem Soc* 76 1804 1954, *R*(+)-isomer: Tattrie et al. *Arch Biochem* 78 319 1958, *Beilstein* 2 IV 1173.] The *racemate* [40290-32-2] is polymorphic with different IR spectra. When crystallised from hexane, or other solvents, the higher melting form with **m** 71.5-72.5° is obtained. The melt then solidifies to give the lower melting  $\alpha$ -form with **m** 49.7-50°. The  $\beta$ -form has **m** 61° (65-66° is also reported). When the lower melting forms are kept at their melting temperatures for a while, they are converted to the higher melting form. [Howe & Malkin *J Chem Soc* 2663 1951, Baer & Kates *J Am Chem Soc* 72 942 1950, *Beilstein* 2 IV 1173.]

***R*-Dipalmitoyl-sn-glycero-3-phosphatidic acid** [7091-44-3] **M 648.9**,  $[\alpha]_D^{26} +4.0$  (c 10, CHCl<sub>3</sub>), **pK<sub>Est(1)</sub> ~1.6**, **pK<sub>Est(2)</sub> ~6.1**. Recrystallise the acid from Me<sub>2</sub>CO at low temperature. At 21° it is soluble in \*C<sub>6</sub>H<sub>6</sub> (4.2%), petroleum ether (0.01%), MeOH (2%), EtOH (2.5%), AcOH (1.3%), Me<sub>2</sub>CO (1.76%), and Et<sub>2</sub>O (1.5%). [Baer *J Biol Chem* 189 235 1951.]

***R*-1,2-Dipalmitoyl-sn-glycero-3-phosphocholine monohydrate (dipalmitoyl- $\alpha$ -L-lecithin)** [63-89-8] **M 752.1**, **m** sinters at 120°,  $[\alpha]_D^{25} +7.0$  (c 5.6, absolute CHCl<sub>3</sub>), **pK<sub>Est</sub> ~5.8 (PO<sub>4</sub>)**. It has three crystalline forms  $\alpha_1$ ,  $\alpha_2$  and  $\beta'$  which change at 60-70° and at 229°, respectively. In order to obtain a fine powder, ~2 g are dissolved in CHCl<sub>3</sub> (15ml) and petroleum ether (b 35-60°) is added; the solution is evaporated to dryness *in vacuo* at <20° and then dried at 0.1mm over CaCl<sub>2</sub>. [Baer & Maurukas *J Am Chem Soc* 74 158 1952, Baer & Kates *J Biol Chem* 185 615 1950, *Beilstein* 4 IV 1463.]

***dl*- $\beta\gamma$ -Dipalmitoylphosphatidyl choline** [2797-68-4] **M 734.1**, **m 230-233°**, **pK<sub>Est</sub> ~5.8 (PO<sub>4</sub>)**. Recrystallise the choline from chloroform and dry it for 48 hours at 10<sup>-5</sup> Torr [O'Leary & Levine *J Phys Chem* 88 1790 1984].

**1,2-Distearoyl-sn-glycerol** [1429-59-0] **M 625.0**. The *dl*-form recrystallises from CHCl<sub>3</sub>/petroleum ether (b 40-60°), **m** 59.5° ( $\alpha$  form) and 71.5-72.5° ( $\beta$  form). Recrystallisation from solvents such as EtOH, MeOH, toluene, Et<sub>2</sub>O gives the higher melting form, and resolidification gives the lower melting form. [IR: Chapman *J Chem Soc* 4680 1958, 2522 1956.] The *S*-isomer is recrystallised from CHCl<sub>3</sub>/petroleum ether and has **m** 76-77°,  $[\alpha]_D^{24} -2.8$  (c 6, CHCl<sub>3</sub>). [Baer & Kates *J Am Chem Soc* 72 942 1950, *Beilstein* 2 IV 1231.]

**1,2-Distearoyl-sn-glycero-3-phosphoethanolamine (distearoyl- $\alpha$ -kephalin)** [1069-79-0] **M 748.1**, **m 180-182° (*R*-form, sintering at 130-135°)**, **m 196° ( $\pm$  form)**, **pK<sub>Est(1)</sub> ~5.8 (PO<sub>4</sub>H)**, **pK<sub>Est(2)</sub> ~10.5 (NH<sub>2</sub>)**. The *R*-form is recrystallised from CHCl<sub>3</sub>/MeOH, and the  $\pm$ -form is recrystallised from EtOH. [Bevan & Malkin *J Chem Soc* 2667 1951, Baer *Can J Biochem Physiol* 81 1758 1959, *Beilstein* 4 IV 1420.]

**Dolichol (from pig liver)** [11029-02-0]  $C_{80}$ - $C_{105}$  **polyprenols**. Crystallise each dolichol 6 times from petroleum ether/EtOH at  $-20^{\circ}$ . The pure individual prenil should run as an entity on a chromatogram on paraffin impregnated paper, with acetone as the mobile phase. [Burgos *Biochem J* **88** 470 1963.]

**Domoic acid [4-(2-carboxyhexa-3,5-dienyl)-3-carboxymethylproline]** [14277-97-5] **M 311.3, m  $215^{\circ}$ ,  $217^{\circ}$ ,  $[\alpha]_D^{20} -108^{\circ}$  (c 1,  $H_2O$ ),  $pK_1$  2.20 ( $2-CO_2H$ ),  $pK_2$  3.72 ( $CO_2H$ ),  $pK_3$  4.93 ( $3-CH_2CO_2H$ ),  $pK_4$  9.82 (NH).** Domoic acid (~300 mg) is purified on a Dowex1 column (3.5 x 40 mm, 200-400 mesh, acetate form), washed with  $H_2O$  until neutral, then eluted with increasing concentrations of AcOH (8L) from 0 to 0.25M. The fraction containing domoic acid (in 50ml) is collected, evaporated to dryness under reduced pressure and recrystallised from aqueous EtOH. It is a glutamate and a Kainate receptor agonist. [Impellizzeri et al. *Phytochemistry* **14** 1549 1975, Takemoto & Diago *Arch Pharm* **293** 627 1960, *Beilstein* **22/4** V 371.]

**Enniatin A** [11113-62-5] **M 681.9, m  $122-122.5^{\circ}$ ,  $[\alpha]_D^{18} -92^{\circ}$  (c 0.9,  $CHCl_3$ ).** It is a cyclic peptidic ester antibiotic which is recrystallised from EtOH/water but is deactivated in alkaline solution. [Ovchinnikov & Ivanov in *The Proteins* (Neurath and Hill eds) Academic Press, NY, Vol V pp. 365 and 516 1982, Tomeda et al. *J Antibiot* **45** 1626 1992.]

**(-)-Ephedrine (1R,2S-2-methylamino-1-phenylpropanol)** [299-42-3] **M 165.2, m  $\sim 34^{\circ}$ ,  $36^{\circ}$ ,  $38.1^{\circ}$ ,  $40^{\circ}$ , b  $126-129^{\circ}/7mm$ ,  $225-227^{\circ}/760mm$ ,  $d^{22} 1.0085$ ,  $[\alpha]_{546}^{20} -47^{\circ}$  and  $[\alpha]_D^{26} -42^{\circ}$  (c 4, 3% HCl),  $[\alpha]_D^{22.5} +15.1^{\circ}$  (c 0.8,  $H_2O$ ),  $-9.36^{\circ}$  (c 3, MeOH),  $pK^{22} 9.58$  ( $pK^{25} 8.84$  in 80% aqueous methoxyethanol).** Purify (-)-ephedrine by vacuum distillation (dehydrates) and forms waxy crystals or granules, and may pick up 0.5  $H_2O$ . The presence of  $H_2O$  raises its melting point to  $40^{\circ}$ . [Moore & Taber *J Amer Pharm Soc* **24** 211 1935.] The *anhydrous base* crystallises from dry ether [Fleming & Saunders *J Chem Soc* 4150 1955]. It gradually decomposes on exposure to light and is best stored in an inert atmosphere in the dark (preferably at  $-20^{\circ}$ ). Its solubility in  $H_2O$  is 5%, in EtOH it is 1% and it is soluble in  $CHCl_3$ ,  $Et_2O$  and mineral oils. It has  $pK_a$  values in  $H_2O$  of 10.25 ( $0^{\circ}$ ) and 8.69 ( $60^{\circ}$ ) [Everett & Hyne *J Chem Soc* 1136 1958, Prelog & Häflinger *Helv Chim Acta* **33** 2021 1950] and  $pK_a^{23} 8.84$  in 80% aqueous methoxyethanol [Simon *Helv Chim Acta* **41** 1835 1958]. The *hydrochloride* has **m  $220^{\circ}$**  (from EtOH/ $Et_2O$ ) and  $[\alpha]_D^{20} -38.8^{\circ}$  (c 2, EtOH). [IR: Chatten & Levi *Anal Chem* **31** 1581 1959.] The *anhydrous base* crystallises from  $Et_2O$  [Fleming & Saunders *J Chem Soc* 4150 1955]. [*Beilstein* **13** H 373, **13**, III 1720, **13** IV 1879.]

**(+)-Ephedrine hydrochloride (1S-2R-2-methylamino-1-phenylpropan-1-ol hydrochloride)** [24221-86-1] **M 201.7, m  $216-219^{\circ}$ ,  $[\alpha]_D^{20} +34^{\circ}$  (c 11.5,  $H_2O$ ).** Recrystallise the hydrochloride from EtOH/ $Et_2O$ . The *free base* crystallises from  $*C_6H_6$  with **m  $40-41^{\circ}$**  (Skita et al. *Chem Ber* **66** 974 1933). [*Beilstein* **13** H 254, **13** II 375, **13** III 1721, **13** IV 1879.]

**(-)-Ephedrine hydrochloride** [50-98-6] **M 201.7, m  $218^{\circ}$ ,  $[\alpha]_{546}^{20} -48^{\circ}$  (c 5, 2M HCl).** It crystallises from water. [*Beilstein* **13** H 254.]

**$\alpha$ -Erythroidine (3R,5S,12S)** [466-80-8] **M 273.3, m  $52-55^{\circ}$ ,  $58-60^{\circ}$ ,  $[\alpha]_D^{27} +136^{\circ}$  (c 0.5,  $H_2O$ ),  $pK_{Est(3)} \sim 7.4$**  Recrystallise  $\alpha$ -erythroidine from pentane. It is best prepared freshly from the more stable hydrochloride. The hydrochloride (1.3g) in  $H_2O$  (20ml) is basified with  $NaHCO_3$  to pH  $\sim 8$  and extracted with  $*C_6H_6$  (6 x 10ml). The combined extracts are evaporated to a small volume and refrigerated. The free base separates as white hygroscopic crystals **m  $52-55^{\circ}$** , which are recrystallised from pentane. Although stable in solution, the crystals turn brown on exposure to air.

**$\alpha$ -Erythroidine hydrochloride** is best purified by dissolving 10g in  $H_2O$  (100ml), adjusting the pH to 8 with aqueous  $NaHCO_3$ , extracting with  $*C_6H_6$  (4 x 20ml), evaporating to 20ml, passing through activated  $Al_2O_3$  and eluting with  $*C_6H_6$ . The eluate is evaporated to a small volume and the crystals are collected, dissolved in EtOH and dry HCl gas passed through to give the pure hydrochloride. When recrystallised from EtOH, it has **m  $226-228^{\circ}(dec)$** , and  $[\alpha]_D^{32} +118^{\circ}$  (c 0.5,  $H_2O$ ). It has  $\lambda_{max}$  at 224nm ( $\epsilon$  35, 500); compare with  $\beta$ -erythroidine hydrochloride below. [Boelkeleide & Grundon *J Am Chem Soc* **75** 2563 1953, Boelkeleide & Morrison *J Am*

*Chem Soc* **80** 3905 1958, abs config: Hill & Shearer *J Org Chem* **27** 3342 1955, *Beilstein* **27** III/IV 3569.]

**$\beta$ -Erythroidine (3*R*,5*S*)** [466-81-9] **M 273.3, m 97-99°, 99.5-100°,  $[\alpha]_D^{25} +89^\circ$  (c 0.5, H<sub>2</sub>O).** Purify it like the  $\alpha$ -isomer but recrystallise it from EtOH. The *free base* is unstable in air and light, but the hydrochloride is more stable and best stored as such. The *methiodide* has **m 211°** (prisms from EtOH).

**$\beta$ -Erythroidine hydrochloride** can be obtained from the  $\alpha$ -isomeric salt as follows: The  $\alpha$ -hydrochloride (1.2g) in 10% aqueous NaOH (12ml) is refluxed for 3 hours under N<sub>2</sub>, cooled in ice and conc HCl added to pH 2. After standing for 3 hours, NaHCO<sub>3</sub> is carefully added to pH 7, the solution is extracted with CHCl<sub>3</sub> (5x), and the extracts are dried, filtered, and concentrated to give an oil which on seeding gives  $\beta$ -erythroidine **m 97-99°**. When dissolved in EtOH and dry HCl gas is passed through the solution, the pure  $\beta$ -erythroidine hydrochloride crystallises out with **m 230.5-231.5°,  $[\alpha]_D^{25} +10^\circ$  (c 0.5, H<sub>2</sub>O).** It is a muscle relaxant with a curare-like-action and is more active than the  $\alpha$ -isomer. [Koniuszy & Folkers *J Am Chem Soc* **72** 519 1950, Boekelheide et al. *J Am Chem Soc* **75** 2550 1953, Boekelheide & Morrison *J Am Chem Soc* **80** 3905 1958, Wenzinger & Boekelheide *J Org Chem* **29** 1307 1964, Berger & Schwartz *J Pharmacol Exp Therap* **93** 362 1948, *Beilstein* **4** IV 3568.]

**Erythromycin A** [114-07-8] **M 733.9, m 133-135°(dec), 135-140°, 137-140°,  $[\alpha]_D^{20} -75^\circ$  (c 2, EtOH), pK<sup>25</sup> 8.9.** It recrystallises from H<sub>2</sub>O to form hydrated crystals which melt at *ca* 135-140°, resolidifies and melts again at 190-193°. The melting point after drying at 56°/8mm is that of the *anhydrous* material and is at 137-140°. Its solubility in H<sub>2</sub>O is ~2mg/ml. The *hydrochloride* has **m 170°, 173°** (from aqueous EtOH, EtOH/Et<sub>2</sub>O). [Flynn et al. *J Am Chem Soc* **76** 3121 1954, constitution: Wiley et al. *J Am Chem Soc* **79** 6062 1957]. [*Beilstein* **18/10** V 398.]

**Ethidium bromide** [1239-45-8] **M 384.3, m 260-262°.** Crystallise it from MeOH or EtOH [Lamos et al. *J Am Chem Soc* **108** 4278 1986]. Its solubility in H<sub>2</sub>O is 1%. [*Beilstein* **22/11** V 352.] **POSSIBLE CARCINOGEN.**

**Ethyl 7-acetoxycoumarin-3-carboxylate** [13207-77-3] **M 276.3, m 147-148°, 152°, 152-154°, pK<sup>25</sup> 7.04.** Purify the ester by recrystallisation from EtOH or aqueous EtOH, dry it *in vacuo* and store it at 2-8°. It is a good fluorogenic substrate for esterases that hydrolyse it to the acid, which acts as an indicator. The fluorescence has  $\lambda_{ex}$  386nm,  $\lambda_{em}$  445nm in 0.1 M Tris buffer pH 8.0. It has been prepared by refluxing 2-imino-3-ethoxycarbonyl-7-hydroxy-2*H*-chromene with Ac<sub>2</sub>O for 1 hour, evaporating and recrystallising the residue from aqueous MeOH [Yasuda & Midorikawa *Bull Chem Soc Jpn* **39** 1754 1966]. *Alternatively*, it can be prepared by acetylation of 7-hydroxy-3-carbethoxycoumarin by the Ac<sub>2</sub>O/pyridine method, and crystallises from EtOH in plates, **m 152°**. An attempted Fries rearrangement by heating the ester with anhydrous AlCl<sub>3</sub> at 16-170° failed, but gave **7-hydroxy-3-carboxycoumarin, m 271° (dec)** (from H<sub>2</sub>O). [Shah & Shah *J Org Chem* **13** 1681 1954, see also Boehm *Archiv der Pharmazie (Weinheim)* **271** 490 1933.]

**Farnesol (trans-trans-3,7,11-trimethyl-2,6,10-dodecatrien-1-ol)** [106-28-5, 4602-84-0 (*trans-trans, E,E*)] **M 222.4, b 111°/0.35mm, 126-127°/0.5mm, 142-143°/2mm,  $d_4^{20}$  0.8871,  $n_D^{20}$  1.4870.** The main impurity is the *cis-trans* isomer. Purify it by gas chromatography using a 4ft x 0.125in 3%OV-1 column at 150°. [Corey et al. *J Am Chem Soc* **92** 6637 1970, Popjak et al. *J Biol Chem* **237** 56 1962.] It has also been fractionated through a 14-in Podbielniak column at 11°/0.35mm. *Alternatively*, it has been purified by gas chromatography using SF96 silicone on Fluoropak columns or Carbowax 20M on Fluoropak or base-washed 30:60 firebrick (to avoid decomposition, prepared by treating the firebrick with 5N NaOH in MeOH and washed with MeOH to pH 8) at 210° with Helium carrier gas at 60 ml/min flow rate. The *diphenylcarbamoyl* derivative has **m 61-63°** (from MeOH) and has an IR band at 3500 cm<sup>-1</sup>. [Bates et al. *J Org Chem* **28** 1086 1963, *Beilstein* **1** IV 2335.]

**Farnesyl acetate (trans-trans-3,7,11-trimethyl-2,6,10-dodecatrien-1-yl acetate)** [4128-17-0 (*trans-trans, E,E*): 29548-30-9 (*isomeric mixture*)] **M 264.4, b 115-125°/0.3mm, 167-169°/0.3mm,  $d_4^{20}$  0.91,  $n_D^{20}$  1.4870.** Purify farnesyl acetate in the same way as the alcohol above. [*Beilstein* **2** I 66, **2** II 154, **2** III 303.]

**Farnesyl pyrophosphate** [13058-04-3, *E,E*: 372-97-4] **M 382.3, pK<sub>Est(1)</sub> ~<2, pK<sub>Est(2)</sub> ~<2, pK<sub>Est(3)</sub> ~3.95, pK<sub>Est(4)</sub> ~6.26.** Purify the pyrophosphate by chromatography on Whatman No3 MM paper in a system of isopropanol-isobutanol/ammonia/water (40:20:1:30) (v/v). Store it as the Li or NH<sub>4</sub> salt at 0°. See geranyl-

geranyl pyrophosphate below.

**Flavin adenine dinucleotide (di-Na, 2H<sub>2</sub>O salt, FAD)** [146-14-5, 84366-81-4 (anhydrous)] **M 865.6**,  $[\alpha]_{546}^{-540}$  (**c 1, H<sub>2</sub>O**). Small quantities of FAD are purified by paper chromatography using *tert*-butyl alcohol/water, cutting out the main spot and eluting with water. Larger amounts can be precipitated from water as the uranyl complex by adding a slight excess of uranyl acetate to a solution at pH 6.0, dropwise and with gentle stirring. The solution is set aside overnight in the cold, and the precipitate is centrifuged off, washed with small portions of cold EtOH, then with cold peroxide-free diethyl ether. It is dried in the dark under vacuum over P<sub>2</sub>O<sub>5</sub> at 50-60°. The uranyl complex is suspended in water, and, after adding sufficient 0.01M NaOH to adjust the pH to 7, the precipitate of uranyl hydroxide is removed by centrifugation [Huennekens & Felton *Methods Enzymol* **3** 954 1957]. It can also be crystallised from water. It should be kept in the dark. More recently it was purified by elution from a DEAE-cellulose (Whatman DE 23) column with 0.1M phosphate buffer pH 7, and the purity was checked by TLC. [Holt & Cotton, *J Am Chem Soc* **109** 1841 1987, *Beilstein* **26** III/IV 3632.]

**Flavin mononucleotide (Na, 2H<sub>2</sub>O salt, riboflavin-5'-phosphate [Na salt, 2H<sub>2</sub>O], FMN)** [130-40-5, 6184-17-4 (Na salt)] **M 514.4**, **pK<sub>1</sub> 2.1 (PO<sub>4</sub>H<sub>2</sub>)**, **pK<sub>2</sub> 6.5 (PO<sub>4</sub>H<sup>-</sup>)**, **pK<sub>3</sub> 10.3 (CONH)**, **fluorescence  $\lambda_{\max}$  530nm (870nm for reduced form)**. Purify FMN by paper chromatography using *tert*-butanol/water, cutting out the main spot and eluting it with water. It can also be purified by adsorption onto an apo-flavodoxin column, followed by elution and freeze drying. It crystallises from aqueous acidic solution. [Mayhew & Strating *Eur J Biochem* **59** 539 1976, *Beilstein* **26** III/IV 2555.]

**5-Fluorouridine (5-fluoro-1- $\beta$ -D-ribofuranosyl-1H-pyrimidine-2,4-dione)** [316-46-1] **M 262.2**, **m 180-182°, 182-184°**,  $[\alpha]_{\text{D}}^{20} +18^{\circ}$  (**c 1, H<sub>2</sub>O**), **pK<sub>Est(1)</sub> ~ 8.0**, **pK<sub>Est(2)</sub> ~ 13**. 5-Fluorouridine is recrystallised from EtOH/Et<sub>2</sub>O and dried at 100° in a vacuum. It has UV with  $\lambda_{\max}$  at 269nm (pH 7.2, H<sub>2</sub>O), 270nm (pH 14, H<sub>2</sub>O). [Liang et al. *Mol Pharmacol* **21** 224 1982, *Beilstein* **24** III/IV 1231.]

**Folic acid (FA, pteroyl-S-glutamic acid)** [59-30-3, 75708-92-8 (2H<sub>2</sub>O)] **M 441.4**, **m >250°(dec)**,  $[\alpha]_{\text{D}}^{25} +23^{\circ}$  (**c 0.5, 0.1N NaOH**), **pK<sub>1</sub> 2.35 (protonation N10)**, **pK<sub>2</sub> 2.75 (protonation N1)**, **pK<sub>3</sub> 3.49 ( $\alpha$ -CO<sub>2</sub>H)**, **pK<sub>4</sub> 4.65 ( $\gamma$ -CO<sub>2</sub>H)**, **pK<sub>5</sub> 8.80 (acidic N3)**. If paper chromatography indicates impurities, then recrystallise it from hot H<sub>2</sub>O or from dilute acid [Walker et al. *J Am Chem Soc* **70** 19 1948]. Impurities may be removed by repeated extraction with *n*-BuOH of a neutral aqueous solution of folic acid (by suspending in H<sub>2</sub>O and adding N NaOH dropwise till the solid dissolves, then adjusting the pH to ~7.0-7.5) followed by precipitation with acid, filtration, or *better* collected by centrifugation and recrystallised from hot H<sub>2</sub>O. [Blakley *Biochem J* **65** 331 1975, Kalifa et al. *Helv Chim Acta* **61** 2739 1978.] Chromatography on cellulose followed by filtration through charcoal has also been used to obtain pure acid. [Sakami & Knowles *Science* **129** 274 1959.] UV:  $\lambda_{\max}$  247 and 296nm ( $\epsilon$  12,800 and 18,700) in H<sub>2</sub>O pH 1.0; 282 and 346nm ( $\epsilon$  27,600 and 7,200) in H<sub>2</sub>O pH 7.0; 256, 284 and 366nm ( $\epsilon$  24600, 24,500 and 86,00) in H<sub>2</sub>O pH 13 [Rabinowitz in *The Enzymes* (Boyer et al. Eds), **2** 185 1960]. [*Beilstein* **26** III/IV 3944.]

**3-Formylchromone (4-oxo-4H-1-benzopyran-3-carbaldehyde)** [17422-74-1] **M 174.2**, **m 151-153°, 152°, 152-153°**. This useful precursor for heterocyclic compounds [Sabitha *Aldrichim Acta* **29** 12 1996] is best prepared by a Vilsmeier-Haack reaction. Tetrachloropyrophosphate (80ml) is added dropwise to a solution of *o*-hydroxyacetophenone (25g, 0.184mole) in DMF (80ml) at -20° to 20° during 10 minutes. The mixture is stirred at ~25° for 13 hours, poured into ice-water and the solid is filtered off, washed with H<sub>2</sub>O, then EtOH, dried *in vacuo*, and recrystallised from Me<sub>2</sub>CO to give white crystals of 3-formylchromone (19.6g, 61%). *Alternatively*, DMF/POCl<sub>3</sub> is used. It has IR bands with  $\nu_{\max}$  at 1605 (CHO), 1650 (C=O) and 1620 (C=C) cm<sup>-1</sup>, the <sup>1</sup>H NMR [60MHz, DMSO-*d*<sub>6</sub>, TMS] has  $\delta$  at 10.16 (s, CHO), 8.9 (s, H-2) and 8.3-7.4 (m, H-5,6,7 and 8), and the MS *m/z*(%) has peaks at 174(4), 147(10), **146(100)**, 121(8), 120(24), 118(3), 105(6), **104(34)**, 92(17), 90(13), 89(11), **76(13)**, 64(12), 63(18), 53(24), fragments in bold are characteristic of the 4-oxo-4H-1-benzopyran-3-carbaldehyde molecule. A large number of substituted derivatives have been prepared in this way and some, particularly the 6-bromoderivative, showed relatively strong anti-anaphylactic activity with low LD<sub>50</sub>, and decreased gastric acid volume. [Nohara et al *Tetrahedron* **30** 3553 1974, Klutchko et al. US Pat 4,098,799 1978, *Chem Abs* **90** 22803 1979.]

**Fructose-1,6-diphosphate (trisodium salt)** [38099-82-0] **M 406.1, pK<sub>1</sub><sup>25</sup> 1.48, pK<sub>2</sub><sup>25</sup> 6.14, pK<sub>3</sub><sup>25</sup> 6.29, pK<sub>4</sub><sup>25</sup> 6.93 (free acid).** Fructose-1,6-diphosphate is best purified *via* the acid strychnine salt which is stable for several months. To remove the strychnine, dissolve 5g in H<sub>2</sub>O (150ml), and add 5N NaOH (or KOH to obtain the K salt) to pH 8.3 (phenolphthalein) with vigorous stirring. Remove the precipitate by centrifugation, wash with cold H<sub>2</sub>O (2x 25ml), extract with CHCl<sub>3</sub> until the extract is free of strychnine (*ca* 8 to 10 times, Mandelein spot test). Freeze-dry the aqueous solution to give the Na salt which is hygroscopic. It has been recrystallised from aqueous EtOH as the *octahydrate*, **m 125°**, **[α]<sub>D</sub><sup>20</sup> +2.6° (c 1, H<sub>2</sub>O).** A neutral solution of the salt keeps well in a frozen state for over several months. [Neuberg et al. *Arch Biochem* **3** 33 1943, Sable *Biochemical Preparations* **2** 52 1952, Stumpf *J Biol Chem* **182** 261 1950]. The *calcium salt* can be partially purified by dissolving in ice-cold N HCl (1g per 10ml) and re-precipitating by dropwise addition of 2N NaOH: the precipitate and supernatant are heated on a boiling water bath for a short time, then filtered, and the precipitate is washed with hot water. The *magnesium salt* can be precipitated from a cold aqueous solution by adding four volumes of EtOH. The *tetramethyl ester* is an oil with **n<sub>D</sub><sup>19</sup> 1.4648, [α]<sub>D</sub><sup>18</sup> +20.2 (c 0.4, MeOH).** [Schulbach & Rauchenberger *Chem Ber* **60** 1178 1927, *Beilstein* **1** IV 4424.]

**Fructose-6-phosphate** [643-13-0] **M 260.1, [α]<sub>D</sub><sup>21</sup> +2.5° (c 3, H<sub>2</sub>O), pK<sub>1</sub><sup>21</sup> 0.97, pK<sub>2</sub><sup>21</sup> 6.11, (pK<sub>2</sub><sup>25</sup> 5.84).** Crystallise fructose-6-phosphate as the barium salt from water by adding 4-volumes of EtOH. The barium can be removed by passage through the H<sup>+</sup> form of a cation exchange resin, and the free acid is collected by freeze-drying. *Alternatively*, the Ba salt is dissolved in H<sub>2</sub>O, and one equivalent of Na<sub>2</sub>SO<sub>4</sub> is added in small portions with stirring, filter off BaSO<sub>4</sub> and freeze dry to give the Na salt. The 6-phosphate hydrolyses more slowly than the 1-phosphate and considerably slower than pyrophosphoric acid (10<sup>2</sup> times) and triphosphoric acid (10<sup>3</sup> times). [Neuberg *Biochem Zeitschrift* **88** 432 1918, pKa: Meyerhof & Lohmann *Biochem Zeitschrift* **185** 113, 131 1927, Neuberg et al. *Arch Biochem* **3** 33, 40 1944, Hydrolysis: Friess *J Am Chem Soc* **74** 5521 1954, *Beilstein* **1** I 464, **1** IV 4423, **H** 31 537.]

**Gangcyclovir** [9-{(1,3-dihydroxy-2-propoxy)methyl}guanine; 2-amino-1,9-{(2-hydroxy-1-hydroxy-methyl)-ethoxymethyl}-6H-purin-6-one; Cytovene; Cymevea(e)n(e)] [82410-32-0] **M 255.2, m >290°(dec), >300°(dec), monohydrate m 248-249°(dec), pK<sub>Est(1)</sub>~1.1, pK<sub>Est(2)</sub>~4.1, pK<sub>Est(3)</sub>~9.7.** Recrystallise gangcyclovir from MeOH. *Alternatively*, dissolve ~90g of it in 700ml of H<sub>2</sub>O, filter and cool (*ca* 94% recovery). It has UV with λ<sub>max</sub> (MeOH) at 254nm (ε 12,880), 270sh nm (ε 9040); its solubility in H<sub>2</sub>O at 25° is 4.3mg/ml at pH 7.0. **ANTIVIRAL.** [Ogilvie et al. *Can J Chem* **60** 3005 1982, Ashton et al. *Biochem Biophys Res Commun* **108** 1716 1982, Martin et al. *J Med Chem* **26** 759 1983.]

**Geraniol** [*trans, E-form* 106-24-1] **M 145.2, b 114-115°/12mm, 229-230°/757mm, d<sub>4</sub><sup>20</sup> 0.889, n<sub>D</sub><sup>20</sup> 1.4628.** Purify geraniol by fractional distillation. It has a sweet rose odour when pure, and the UV has λ<sub>max</sub> at 190-195nm (ε 18,000). [See also p 146, *Beilstein* **1** H 457, **1** IV 2277.]

**Geranyl acetate** [*trans, E-form* 105-87-3; 16409-44-2] **M 196.3, b 118°/12mm, 138°/25mm, 236-242°/760mm(dec), d<sub>15</sub><sup>15</sup> 0.9174, n<sub>D</sub><sup>15</sup> 1.4766.** Purify the fragrant smelling geranyl acetate by fractional distillation at as high a vacuum as possible. It is very soluble in EtOH but insoluble in H<sub>2</sub>O. [*Beilstein* **2** H 140, **2** I 65, **2** II 153, **2** III 299, **2** IV 204.]

**Geranylgeranyl pyrophosphate** [6699-20-3 (NH<sub>4</sub> salt), 64732-91-8] **M 450.5, pK<sub>Est(1)</sub>~<2, pK<sub>Est(2)</sub>~<2, pK<sub>Est(3)</sub>~3.95, pK<sub>Est(4)</sub>~6.26.** Purify the pyrophosphate by countercurrent distribution between two phases of a butanol/isopropyl ether/ammonia/water mixture (15:5:1:19) (v/v), or by chromatography on DEAE-cellulose (linear gradient of 0.02M KCl in 1mM Tris buffer, pH 8.9). *Alternatively*, purify it through a column of Dowex 1-x8 (formate form previously washed with MeOH) and eluted with a linear gradient of 0.053—0.43M ammonium formate in a total volume of 300ml of MeOH. The purity can be checked by TLC on Silica gel G on buffered plates (pH 6.5), eluted with CHCl<sub>3</sub>/MeOH/H<sub>2</sub>O (60:40:9) and developed with I<sub>2</sub> vapour. Store it as a powder at 0°. [Altman et al. *J Am Chem Soc* **94** 3257 1972.] It is more stable as the di(tri-n-butylammonium)hydrogen phosphate salt which can be obtained from the acid by evaporation in a rotary evaporator below 32° [Upper & West *J Biol Chem* **242** 3285 1967]. [Gregonis & Rilling *Biochemistry* **13** 1538 1974, Gregonis & Rilling

*Biochem Biophys Res Commun* **54** 449 1973.]

**Geranyl pyrophosphate** [*E*-form 763-10-0 tri-(NH<sub>4</sub> salt), *Z*-form 16751-02-3] **M 314.2**, **pK<sub>Est(1)</sub> ~ 2**, **pK<sub>Est(2)</sub> ~ 2**, **pK<sub>Est(3)</sub> ~ 3.95**, **pK<sub>Est(4)</sub> ~ 6.26**. Purify the pyrophosphate by paper chromatography on Whatman No 3 MM paper in a system of isopropyl alcohol/isobutyl alcohol/ammonia/water (40:20:1:39), R<sub>F</sub> 0.77-0.82. Store it in the dark as the ammonium salt at 0°. The *E*-form crystallises in platelets from aqueous Me<sub>2</sub>CO, **m ~ 120°**. It dissolves in dry MeCN. *Alternatively*, purify it through a column of Dowex AG 1x8(200-400mesh) equilibrated with 50mM NH<sub>4</sub> formate, and elute with MeOH/H<sub>2</sub>O/NH<sub>4</sub>OH (95:5:05), then freeze-dry. [Dixit et al. *J Org Chem* **46** 1967 1981, *Beilstein* **1** IV 3580.]

**Gliotoxin (3*R*-6*t*-hydroxy-3-hydroxymethyl-2-methyl-(5*at*)-2,3,6,10-tetrahydro-5*aH*-3,10*ac*-epidissulfido-[1,2-*a*]-indol-1,4-dione)** [67-99-2] **M 326.4**, **m 191-218°(dec)**, **220°(dec)**, **221°(dec)**, **[α]<sub>D</sub><sup>20</sup> -254° (c 0.6, CHCl<sub>3</sub>)**, **[α]<sub>D</sub><sup>25</sup> -270° (c 1.7, pyridine)**. Purify gliotoxin by recrystallisation from MeOH. Its solubility in CHCl<sub>3</sub> is 1%. The *dibenzoyl* derivative has **m 202°** (from CHCl<sub>3</sub>/MeOH). [Glister & Williams *Nature* **153** 651 1944, Elvidge & Spring *J Chem Soc Suppl* **135** 1949, Johnson et al. *J Am Chem Soc* **65** 2005 1943, Bracken & Raistrick *Biochem J* **41** 569 1947.]

**Glucose-1-phosphate (G-1-P)** [59-56-3] **M 260.1**, **[α]<sub>D</sub><sup>25</sup> +120° (c 3, H<sub>2</sub>O)**, **[α]<sub>D</sub><sup>20</sup> +78° (c 4, H<sub>2</sub>O of di-K salt)**, **pK<sub>1</sub> 1.11**, **pK<sub>2</sub> 6.13** [**pK<sub>25</sub> 6.50**]. Two litres of a 5% aqueous solution of the phosphate are purified by adjusting the pH to 3.5 with glacial acetic acid (+ 3g of charcoal) and filtering. An equal volume of EtOH is added, the pH is adjusted to 8.0 (glass electrode) and the solution is stored at 3° overnight. The precipitate is filtered off, dissolved in 1.2L of distilled water, filtered and an equal volume of EtOH is added. After standing at 0° overnight, the crystals are collected at the centrifuge and washed with 95% EtOH, then absolute EtOH, ethanol/diethyl ether (1:1), and diethyl ether. [Sutherland & Wosilait, *J Biol Chem* **218** 459 1956.] Its barium salt can be crystallised from water and EtOH. Heavy metal impurities are removed by passage of an aqueous solution (*ca* 1%) through an Amberlite IR-120 column (in the appropriate H<sup>+</sup>, Na<sup>+</sup> or K<sup>+</sup> forms). *Di-K salt* crystallises as a *dihydrate* from EtOH. [see McGready *Biochemical Preparations* **4** 63 1955.] [*Beilstein* **17/8** V 247.]

**Glucose-6-phosphate (G-6-P)** [acid 156-73-5; Ba salt 58823-95-3; Na salt 54010-71-8] **M 260.1**, **m 205-207°(dec)** **mono Na salt**, **[α]<sub>D</sub><sup>20</sup> +41° (c 5, H<sub>2</sub>O)**, **pK<sub>1</sub> 1.65**, **pK<sub>2</sub> 6.11**, **pK<sub>3</sub><sup>25</sup> 11.71** [-C<sub>1</sub>(OH)O<sup>-</sup>]. It can be freed from metal impurities as described for glucose-1-phosphate. The solubility of the Na salt is 5% in H<sub>2</sub>O at 20°. Its *barium* salt can be purified by solution in dilute HCl and precipitation by neutralising the solution. The precipitate is washed with small volumes of cold water and dried in air. *Alternatively*, the barium salt is dissolved in H<sub>2</sub>O, 4 volumes of EtOH are added, the precipitate is collected, washed with 90% EtOH, absolute EtOH, 75% EtOH/25% Et<sub>2</sub>O, 25% EtOH/75% Et<sub>2</sub>O and finally dry Et<sub>2</sub>O. The dry Barium salt has **[α]<sub>D</sub><sup>25</sup> +17.9° (c 1, H<sub>2</sub>O)**. [*Beilstein* **1** IV.]

G-6-P is relatively more stable to hydrolysis (12% hydrolysis in N HCl/100° in 4 hours) than G-1-P (45% hydrolysis in N HCl/20° in 20 hours). [Lardy & Fischer *Biochemical Preparations* **2** 39 1952.]

**L-α-Glycerol phosphocholine (Cadmium Chloride)<sub>x</sub> complex** [64681-08-9] **M 257.2 + (183.3)<sub>x</sub>**, **pK<sub>Est</sub> ~ 5.5**. Glycerol phosphocholine is purified *via* the CdCl<sub>2</sub> complex which is recrystallised four times from 99% EtOH by standing at 0° for 1 hour. The white precipitate is collected, washed with EtOH, Et<sub>2</sub>O and dried in a vacuum. The amorphous Cd complex can be converted to the crystalline form [C<sub>8</sub>H<sub>20</sub>O<sub>6</sub>NP.CdCl<sub>2</sub>.3H<sub>2</sub>O] by dissolving 34.4g in H<sub>2</sub>O (410ml), and 99% EtOH (1650ml total) is added slowly with stirring and allowing the clear solution to stand at 25° for 12 hours, then at 5° for 12 hours. The crystalline complex is filtered off, washed with cold 80% EtOH and dried in air. Glycerol phosphocholine can be recovered from the complex by dissolving it in H<sub>2</sub>O (2% solution) and passing it through an ion-exchange column (4.9 x 100cm, of 1 volume IRC-50 and 2 volumes of IR-45). The effluent is concentrated to a thick syrup at 45°. It is dried further at 50°/P<sub>2</sub>O<sub>5</sub>/48 hours. The vitreous product (~8.25g) is then dissolved in 99% EtOH (50ml), and the clear solution is cooled to 5°, whereby crystals begin to appear, and crystallisation is completed at -15° after 16 hours. The crystals are filtered off, washed with 99% EtOH, and Et<sub>2</sub>O then dried at 50° *in vacuo* over P<sub>2</sub>O<sub>5</sub>. It can be recrystallised from 99.5% EtOH (long prisms). It is *hygroscopic* and must be handled in a H<sub>2</sub>O-free atmosphere [Tattie & McArthur *Biochemical Preparations* **6** 16 1958, Baer & Kates *J Am Chem Soc* **70** 1394 1948, *Acta*

*Cryst* 21 79, 87 1966].

**Hematin (ferrihaeme hydroxide)** [15489-90-4] **M 633.5, m 200°(dec), pK<sub>Est</sub> ~ 4.** Crystallise it from pyridine. Dry it at 40° *in vacuo*. [Beilstein 26 III/IV 3047.]

**Hematoporphyrin IX (8,13-bis(1-hydroxyethyl)-3,7,12,17-tetramethyl-21H-23H-porphin-2,18-dipropionic acid, 3,3'-[7,12-bis-(1-hydroxyethyl)-3,8,13,17-tetramethyl-porphyrin-2,18-diyl]-dipropionic acid)** [14459-29-1] **M 598.7, pK<sub>Est</sub> ~4.8.** Purify it by dissolving it in EtOH and adding H<sub>2</sub>O or Et<sub>2</sub>O to give deep red crystals. It has also been recrystallised from MeOH. Its UV has  $\lambda_{\text{max}}$  at 615.5, 565, 534.4 and 499.5nm in 0.1N NaOH, and 597, 619, 634, 653, 683 and 701nm in 2N HCl. [Falk *Porphyryns and Metalloporphyryns* Elsevier, NY, p 175 1964, LCCCN<sub>o</sub> 63-19821.] It is used in the affinity chromatographic purification of Heme proteins [Olsen *Methods Enzymol* 123 324 1986]. The *O*-methyl-dimethyl ester has **m 203-206°** (from CHCl<sub>3</sub>/MeOH), and the *O,O'*-dimethyl-dimethyl ester has **m 145°** (from CHCl<sub>3</sub>/MeOH). [Paul *Acta Chem Scand* 5 389 1951, Beilstein 26 III/IV 3157, and 3158 for the HCl.]

**Hematoporphyrin dimethyl ester** [33070-12-1] **M 626.7, m 212°.** It crystallises from CHCl<sub>3</sub>/MeOH. [Beilstein 26 III/IV 3157.]

**Hematoxylin ( $\pm$ -11bc-7,11b-dihydroindeno[2,1-c]-chromen-3,4-6ar-9,10-pentaol)** [517-28-2] **M 302.3, m 200°(dec), 210-212°(dec).** Hematoxylin recrystallises from H<sub>2</sub>O (as *trihydrate*) in white-yellow crystals which become red on exposure to light and then melt at 100-120°. It has been recrystallised from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>. It has been recrystallised as well from dilute aqueous NaHSO<sub>3</sub> until colourless and is soluble in alkali, borax and glycerol. Store it in the dark below 0°. [Morsingh & Robinson *Tetrahedron* 26 182 1970, Dann & Hofmann *Chem Ber* 98 1498 1955, Beilstein 17/8 V 469.]

**Hemin (ferriprotoporphyrin IX chloride)** [16009-13-5] **M 652.0, m sinters at 240°, pK<sub>Est</sub> ~4.8.** Hemin is purified by recrystallisation from AcOH. Also, hemin (5g) is shaken in pyridine (25ml) till it dissolves, then CHCl<sub>3</sub> (40ml) is added, the container is stoppered and shaken for 5 minutes (releasing the stopper occasionally). The solution is filtered under slight suction, and the flask and filter are washed with a little CHCl<sub>3</sub> (15ml). During this period, AcOH (300ml) is heated to boiling, and saturated aqueous NaCl (5ml) and conc HCl (4ml) are added. The CHCl<sub>3</sub> filtrate is poured in a steady stream, with stirring, into the hot AcOH mixture and set aside for 12 hours. The crystals are filtered off, washed with 50% aqueous AcOH (50ml), H<sub>2</sub>O (100ml), EtOH (25ml), Et<sub>2</sub>O and dried in air. [Fischer *Org Synth Coll Vol III* 442 1955, Beilstein 26 III/IV 3048.]

**(+)-Hydroquinidine anhydrous (9S-6'-methoxy-10,11-dihydrocinchonon-9-ol)** [1435-55-8] **M 326.4, m 168-169°, 169°, 169-170°, 171-172°, [α]<sub>D</sub><sup>20</sup> +231° (c 2, EtOH), +299° (c 0.82, 0.1N H<sub>2</sub>SO<sub>4</sub>), pK<sub>Est</sub> ~ 8.8.** (+)-Hydroquinidine forms needles from EtOH and plates from Et<sub>2</sub>O. It is slightly soluble in Et<sub>2</sub>O and H<sub>2</sub>O but readily soluble in hot EtOH. [Heidelberger & Jacobs *J Am Chem Soc* 41 826 1919, King *J Chem Soc* 523 1946.] The *hydrochloride* has **m 273-274°, [α]<sub>D</sub><sup>26</sup> +184° (c 1.3, MeOH)** and is very soluble in MeOH and CHCl<sub>3</sub>, but less soluble in H<sub>2</sub>O, EtOH and still less soluble in dry Me<sub>2</sub>CO. [Kyker & Lewis *J Biol Chem* 157 707 1945, Emde *Helv Chim Acta* 15 557 1932, Beilstein 23/13 V 340.]

**Hydroquinine** [522-66-7] **M 326.4, m 168-171°, 171.5°, [α]<sub>D</sub><sup>16</sup>-143° (c 1.087, EtOH), pK<sub>15</sub> 5.33 and 8.87.** Recrystallise hydroquinine from EtOH, Et<sub>2</sub>O or \*C<sub>6</sub>H<sub>6</sub>. [Rabe & Schultz *Chem Ber* 66 120 1933, Beilstein 23 III/IV 3193, 23/13 V 340.] The *racemate* has **m 175-175.5°** (from EtOH).

**5-Hydroxycreatinine (creatol, CTL, 2-amino-5-hydroxy-1-methylimidazolidin-4-one)** [133882-98-1] **M 129.1, m 190° (191°, dec) for the hydrochloride, pK<sub>25</sub> 4.2.** Creatol is one of three metabolites of creatinine [60-27-5] which was isolated from inflamed rabbit skin tissues [Ienaga et al. *Tetrahedron Lett* 28 4587 1987], and the urine of uraemic rats (adenine induced chronic renal failure) and uraemic patients. The other two metabolites are 1-methylhydantoin (1-methylimidazolidine-2,4-dione [616-04-6]) and 5-hydroxy-1-methylhydantoin [1-methyl-5-hydroxyimidazolidine-2,4-dione] [84210-26-4]. [Ienaga et al. *J C S, Chem Commun*

509 1991.] Urine from uraemic rats (44ml, collected after 24 hours of administration of creatinine) was applied onto a PK216 ion exchange column (H<sup>+</sup> form) that was washed with H<sub>2</sub>O, then eluted with aqueous NH<sub>3</sub> (2mol dm<sup>-3</sup>) and evaporated *in vacuo*. The residue, in H<sub>2</sub>O, was applied onto a Biolex-70 ion exchange column (H<sup>+</sup> form) that was washed with H<sub>2</sub>O, then eluted with aqueous AcOH (0.1mol dm<sup>-3</sup>), and the eluate was evaporated *in vacuo* to give *creatol* (0.66mg, 0.11%, purity checked by ODS reverse phase HPLC). [The H<sub>2</sub>O wash from the PK216 column gave the other two metabolites (1.5mg, 0.35 and 0.87mg, 0.15% respectively) which were separated on a silica gel column with EtOAc elution.] Creatol has <sup>1</sup>H NMR (D<sub>2</sub>O, *tert*-BuOH as internal standard) with δ at 3.11 (s, in N<sup>12</sup>CH<sub>3</sub> creatol; but d, in N<sup>13</sup>CH<sub>3</sub> labelled creatol with *J* = 143 Hz) and 5.11 (s, H-5); and <sup>13</sup>C NMR (D<sub>2</sub>O, dioxane as internal standard) with δ at 30 (NCH<sub>3</sub>), 81 (C-5), 160 (C-2) and 174 (C-4). [Ienaga et al. *J C S, Chem Commun* 509 1991.] The natural enantiomer most probably has the *S* configuration at C-5 [see Ienaga et al. *J C S, Chem Commun* 509 1991]. [Nakamura & Ienaga, Glycoyamidine derivatives, *Japanese Kokai Tokkyo Koho*, JP 2957217 1989, US patent 4957936 1990.] Data kindly supplied by Dr Kazu Ienaga, Nippon Zoki Pharmaceutical Co Ltd, Osaka, Japan.

***R*-(*-*)-2-Hydroxy-3,3-dimethyl-γ-butyrolactone (3-hydroxy-4,4-dimethyl-4,5-dihydrofuran-2-one, D-pantolactone) [599-04-2] M 130.1, m 89-91°, 90.5-91.5°, 91°, 92-93°, b 120-122°/15mm, [α]<sub>D</sub><sup>20</sup> -28° (c 5, MeOH), [α]<sub>D</sub><sup>20</sup> -51° (c 3, H<sub>2</sub>O). Recrystallise the lactone from Et<sub>2</sub>O/petroleum ether, diisopropyl ether or \*C<sub>6</sub>H<sub>6</sub>/petroleum ether and sublime it at 25°/0.0001mm. It hydrolyses readily to the hydroxy-acid and racemises when heated above 145°. The *Brucine salt* has m 211-212° (from EtOH). [Kuhn & Wieland *Chem Ber* 73 1134 1940, and Stiller et al. *J Am Chem Soc* 62 1779 1940, Bental & Tishler *J Am Chem Soc* 68 1463 1946, *Beilstein* 18/1 V 22.]**

(±)-5-Hydroxy-1-methyl-imidazolidine-2,4-dione (HMH) [84210-26-4] M 130.1, m 138°, pK 8.64. The racemic compound was obtained by oxidation of *N*-1-methylhydantoin with Pb(OAc)<sub>4</sub> in \*C<sub>6</sub>H<sub>6</sub> (60°, 24 hours), isolation of the 5-acetoxy derivative (*via* a short silica-gel column chromatography; eluting with \*C<sub>6</sub>H<sub>6</sub>/hexane), followed by hydrolysis with aqueous 0.3M H<sub>2</sub>SO<sub>4</sub> (~25°, 40 minutes). Neutralisation, followed by evaporation, and silica-gel TLC (EtOAc) provided pure racemate m 138°, (64% yield) with the expected <sup>1</sup>H NMR spectrum and EI MS and elemental analysis. It is a plant growth regulator affecting the flowering period of cut white chrysanthemums at concentrations as low as 10<sup>-6</sup> M. [Ienaga et al. *Tetrahedron Lett* 28 4587 1987.]

The natural metabolite *N*-1-methylhydantoin (MH [616-04-6], see below) was isolated from rabbit urine by diluting the urine two-fold with MeOH and evaporating. The residue was extracted with MeOH/EtOAc (1:6), filtered, the filtrate was concentrated, subjected to silica-gel column chromatography and eluted with MeOH/CHCl<sub>3</sub> (1:9). The first eluate gave the pure crystalline 5-hydroxy compound (HMH, 28% of total) that had m 147-150° and [α]<sub>D</sub><sup>25</sup> -5.1° (c 1.0, MeOH). The expected <sup>1</sup>H NMR [Eu(tfc)<sub>3</sub>] spectrum indicated that its enantiomeric excess was 56% (i.e. a 78:22 mixture). Some racemisation of the 5-hydroxy compound may well have occurred during isolation. The synthetic racemate was resolved by conversion into the diastereoisomeric Bz-*L*-proline ester (1:1, *R*:*S*) [with 1-(3-diethylaminopropyl)-3-ethylcarbodiimide HCl in the presence of DMAP in MeCN at 0°/2 hours) followed by purification through a silica-gel column and eluting with EtOAc/hexane (7:3)] in 56% yield. Fractional recrystallisation from Me<sub>2</sub>CO gave pure 5-(*R*)-(*N*-Bz-*L*-prolyloxy)-1-methylimidazolidine-2,4-dione (30%), m 189-192°, [α]<sub>D</sub><sup>25</sup> -2.1° (c 1.0, MeOH) {<sup>1</sup>H NMR, DMSO-*d*<sub>6</sub> has δ at 1.8-2.0 (m, 3H), 2.2-2.4 (m, 1 H), [2.52 (s) + 2.75 (s), 3H], 3.4-3.55 (m, 2 H), [4.35 (dd, *J* = 4, 9 Hz) + 4.46 (dd, *J* = 4, 9 Hz) 1H], [5.03 (s) + 5.10 (s) 2H], [6.90 (s) + 6.10 (s) 1H], 7.25-7.45 (m, 5H, aromatic) and 11.21 (br s, NH)}; and when these crystals were subjected to X-ray crystallography they revealed that the absolute configuration at C5 of the *imidazolidine-2,4-dione* part of the molecule was *R*, and that C5 of the *natural metabolite* was therefore *S* when a comparison of the HPLC of the mixture of prolyl-esters from the natural metabolite was made. The residue from evaporation of the mother liquors was recrystallised from EtOAc and gave the diastereoisomeric 5-(*S*)-*N*-Bz-*L*-prolyl ester (8%) with m 147-150°, [α]<sub>D</sub><sup>25</sup> -68.9° (c 1.0, MeOH) {<sup>1</sup>H NMR, DMSO-*d*<sub>6</sub> has δ at 1.7-2.0 (m, 3H), 2.2-2.4 (m, 1 H), [2.63 (s) + 2.73(s), 3H], 3.35-3.55 (m, 2 H), [4.36 (dd, *J* = 4, 9 Hz) + 4.47 (dd, *J* = 4, 9 Hz) 1H], [5.04 (d, *J* = 13Hz) + 5.07 (d, *J* = 13Hz) 1H], [5.07 (d, *J* = 13Hz) + 5.13 (d, *J* = 13Hz) 1H], [6.05 (s) + 6.12 (s) 1H], 7.25-7.45 (m, 5H, aromatic) and 11.25(br s, NH)}. Attempted hydrolysis of these esters gave racemic 5-hydroxyimidazolidine-2,4-dione, however, owing to facile racemisation. [Ienaga et al. *J Chem Soc, Perkin Trans I* 1153 1989.] The racemate and the optically active form of 5-hydroxy-1-methylhydantoin have the same <sup>1</sup>H NMR (400MHz, Me<sub>2</sub>CO-*d*<sub>6</sub>, TMS) with δ at ~5.1 (d, C-5) and ~5.9 (d, 1-NMe). Data kindly supplied by Dr Kazu Ienaga, Nippon Zoki Pharmaceutical Co Ltd, Osaka, Japan.

**(-)-Inosine** [58-63-9] M 268.2, m 90° (dihydrate), 218° (dec, anhydrous),  $[\alpha]_{546}^{20}$  -76° (c 1, 0.1M NaOH),  $pK_1^{25}$  1.06,  $pK_2^{25}$  8.96,  $pK_3^{25}$  11.36. (-)-Inosine forms *anhydrous* crystals from aqueous 80% EtOH but the *dihydrate* from H<sub>2</sub>O. [Beilstein 31 H 25, 26 III/IV 2087.]

**myo-Inositol** (cyclohexane[1*r*,2*c*,3*c*,4*t*,5*c*,6*t*]-hexol) [87-89-8] M 180.2, m 218° (di-hydrate), 225-227°, 226-230°. Recrystallise *myo*-inositol from aqueous 50% ethanol or H<sub>2</sub>O forming a *dihydrate*, or *anhydrous* crystals from AcOH. The dihydrate is efflorescent and becomes anhydrous when heated at 100°. The anhydrous crystals are not hygroscopic. Its solubility in H<sub>2</sub>O at 25° is 14%, at 60° it is 28%; it is slightly soluble in EtOH but insoluble in Et<sub>2</sub>O. [Ballou & Anderson *J Am Chem Soc* 75 748 1953, Anderson & Wallis *J Am Chem Soc* 70 2931 1948, Beilstein 6 II 1157, 6 IV 7919.]

**Inositol monophosphate** [15421-51-9] M 260.1, m 195-197°(dec). Crystallise the phosphate from water, and EtOH. Recrystallise 1g by dissolving it in 3ml of H<sub>2</sub>O and adding slowly 15ml of commercial EtOH, filter the crystals, wash with a little EtOH then Et<sub>2</sub>O and dry it in a vacuum. [McCormick & Carter *Biochemical Preparations* 2 65 1952.]

**5-Iodouridine** (5-iodo-1-[β-D-ribofuranosyl]-pyrimidine-2,4(1*H*)-dione) [1024-99-3] M 370.1, m 205-208°(dec), 210-215°(dec),  $[\alpha]_D^{20}$  -23.5° (c 1, H<sub>2</sub>O),  $pK^{20}$  8.5. Recrystallise 5-iodouridine from H<sub>2</sub>O and dry it *in vacuo* at 100°. It UV has  $\lambda_{max}$  at 289nm (0.01N HCl) and 278nm (0.01N NaOH). [Prusoff et al. *Cancer Res* 13 221 1953, Beilstein 24 III/IV 1235.]

**Isopentenyl pyrophosphate** [358-71-4] M 366.2,  $pK_{Est(1)} \sim 2$ ,  $pK_{Est(2)} \sim 2$ ,  $pK_{Est(3)} \sim 3.95$ ,  $pK_{Est(4)} \sim 6.26$ . Purify the pyrophosphate by chromatography on Whatman No 1 paper using *tert*-butyl alcohol/formic acid/water (20:5:8, R<sub>F</sub> 0.60) or 1-propanol/ammonia/water (6:3:1, R<sub>F</sub> 0.48). *Alternatively*, purify it by chromatography on a DEAE-cellulose column or a Dowex-1 (formate form) ion-exchanger using formic acid and ammonium formate as eluents. A further purification step is to convert it to the *monocyclohexylammonium salt* by passage through a column of Dowex-50 (cyclohexylammonium form) ion-exchange resin. It can also be converted into its lithium salt. (See geranyl pyrophosphate above.)

**Kanamycin B** (Bekanamycin, 4-*O*-[2,6-diamino-2,6-dideoxy-α-D-glucopyranosyl]-6-*O*-[3-amino-3-deoxy-α-D-glucopyranosyl]-2-deoxystreptamine) [4696-76-8, 29701-07-3 (sulfate salt)] M 483.5, m 170-179°(dec), 178-182°(dec),  $[\alpha]_D^{18}$  +130° (c 0.5, H<sub>2</sub>O),  $pK^{25}$  7.2. A small quantity of kanamycin B (24mg) can be purified on a small Dowex-1 x 2 column (6 x 50mm); the required fraction is evaporated to dryness and the residue crystallised from EtOH containing a small amount of H<sub>2</sub>O. [Umezawa et al. *Bull Chem Soc Jpn* 42 537 1969.] It has been crystallised from H<sub>2</sub>O by dissolving ~1g in H<sub>2</sub>O (3ml), adding Me<sub>2</sub>NCHO (3ml) and setting aside at 4° overnight. The needles are collected and dried to constant weight at 130°. It has also been recrystallised from aqueous EtOH. It is slightly soluble in CHCl<sub>3</sub> and isoPrOH. [IR: Wakazawa et al. *J Antibiot* 14A 180, 187 1961, Ito et al. *J Antibiot* 17 A 189 1964, Beilstein 18 III/IV 7631.]

**Lecithin** (1,2-diacylphosphatidylcholine mixture) [8002-43-5] M ~600-800, **amorphous**. Lecithin from hen egg white is purified by solvent extraction and chromatography on alumina. It is suspended in H<sub>2</sub>O and kept frozen until required [Lee & Hunt *J Am Chem Soc* 106 7411 1984, Singleton et al. *J Am Oil Chem Soc* 42 53 1965]. For purification of commercial egg lecithin, see Pangborn [*J Biol Chem* 188 471 1951].

**Leucopterin** (2-amino-5,8-dihydropteridine-4,6,7(1*H*)-trione) [492-11-5] M 195.1, m >300° (dec),  $pK_1^{20}$  -1.66,  $pK_2^{20}$  7.56,  $pK_3^{20}$  9.78,  $pK_4^{20}$  13.6. Leucopterin is purified by dissolving it in aqueous NaOH, stirring with charcoal, filtering and precipitating by adding aqueous HCl, then drying at 100° in a vacuum. It separates with 0.5 mole of H<sub>2</sub>O. Its solubility in H<sub>2</sub>O is 1g/750 litres [Albert et al. *J Chem Soc* 4219 1952, Albert & Wood *J Appl Chem (London)* 2 591 1952, Pfeleiderer *Chem Ber* 90 2631 1957]. [Beilstein 26 III/IV 4017.]

**DL-α-Lipoamide** (±-6,8-thioctic acid amide, 5-[1,2]-dithiolan-3-ylvaleric acid amide) [940-69-2; 3206-73-

3] **M 205.3, m 124-126°, 126-129°, 130-131°.** DL- $\alpha$ -Lipoamide is recrystallised from EtOH and has UV with  $\lambda_{\max}$  at 331nm in MeOH. [Reed et al. *J Biol Chem* **232** 143 1958, IR: Wagner et al. *J Am Chem Soc* **78** 5079 1956, *Beilstein* **19/7** V 238.]

**DL- $\alpha$ -Lipoic acid ( $\pm$ -6,8-thioctic acid, 5-[1,2]-dithiolan-3-ylvaleric acid)** [1077-28-7] **M 206.3, m 59-61°, 60.5-61.5° and 62-63°, b 90°/10<sup>-4</sup>mm, 150°/0.1mm, pK<sup>25</sup> 4.7.** It forms yellow needles from cyclohexane or hexane and has been distilled at high vacuum: and sublimes at  $\sim$ 90° and very high vacuum. It is insoluble in H<sub>2</sub>O but dissolves in alkaline solution. [Lewis & Raphael *J Chem Soc* 4263 1962, Soper et al. *J Am Chem Soc* **76** 4109, Reed & Niu *J Am Chem Soc* **77** 416 1955, Tsuji et al. *J Org Chem* **43** 3606 1978, Calvin *Fed Proc USA* **13** 703 1954.] The *S*-benzylisothiuronium salt has **m** 153-154° (evacuated capillary, from MeOH), 132-134°, 135-137° (from EtOH). The *d*- and *l*- forms have **m** 45-47.5° and  $[\alpha]_{\text{D}}^{25} \pm 113^{\circ}$  (c 1.88, \*C<sub>6</sub>H<sub>6</sub>) and have UV in MeOH with  $\lambda_{\max}$  at 330nm ( $\epsilon$  140). [*Beilstein* **19/7** V 237.] The reduced form, ( $\pm$ )-6,8-dimercaptooctanoic acid, [7516-48-5] **M 208.3**, is a light yellow liquid which is sold in sealed ampoules.

**D-Luciferin (firefly luciferin, S-2[6-hydroxybenzothiazol-2-yl]-4,5-dihydrothiazol-4-carboxylic acid)**, [2591-17-5] **M 280.3, m 189.5-190°(dec), 196°(dec), 201-204°, 205-210°(dec, browning at 170°),  $[\alpha]_{\text{D}}^{22}$  -36° (c 1.2, DMF), pK<sub>Est(1)}<sup>~</sup> 1.2 (benzothiazole-N), pK<sub>Est(2)}<sup>~</sup> 1.6 (thiazolidine-N), pK<sub>Est(3)}<sup>~</sup> 6.0 (CO<sub>2</sub>H), pK<sub>Est(4)} 8.5 (6OH).</sub></sub></sub></sub>** D-Luciferin crystallises as pale yellow needles from H<sub>2</sub>O, or MeOH (83mg/7ml). It has UV with  $\lambda_{\max}$  at 263 and 327nm (log  $\epsilon$  3.88 and 4.27) in 95% EtOH. The Na salt has a solubility of 4mg in 1 ml of 0.05M glycine. [White et al. *J Am Chem Soc* **83** 2402 1961, **85** 337 1963, UV and IR: Bitler & McElroy *Arch Biochem* **72** 358 1957, Review: Cormier et al. *Fortschr Chem Org Naturst* **30** 1 1973, *Beilstein* **27** III/IV 8934.]

**Lumiflavin (7,8,10-trimethylbenzo[g]pteridine-2,4(3H,10H)-dione)** [1088-56-8] **M 256.3, m 330°(dec), 340°(dec), pK<sup>25</sup> 10.2.** Lumiflavin forms orange crystals upon recrystallisation from 12% aqueous AcOH, or from formic acid. It sublimes at high vacuum. It is freely soluble in CHCl<sub>3</sub>, but not very soluble in H<sub>2</sub>O and most organic solvents. In H<sub>2</sub>O and CHCl<sub>3</sub> solution it has a green fluorescence. Its UV has  $\lambda_{\max}$  at 269, 355 and 445nm ( $\epsilon$  38,800, 11,700 and 11,800, respectively) in 0.1N NaOH and 264, 373 and 440nm ( $\epsilon$  34,700, 11,400 and 10,400, respectively) in 0.1N HCl, while the UV in CHCl<sub>3</sub> has  $\lambda_{\max}$  at 270, 312, 341, 360, 420, 445 and 470nm. [Hemmerich et al. *Helv Chim Acta* **39** 1242 1956, Holiday & Stern *Chem Ber* **67** 1352 1834, Yoneda et al. *Chem Pharm Bull Jpn* **20** 1832 1972, Birch & Moyer *J Chem Soc* 2622 1958, Fluorescence: Kuhn & Moruzzi *Chem Ber* **67** 888 1934, *Beilstein* **26** III/IV 2539.]

**Magnesium protoporphyrin dimethyl ester** [14724-63-1] **M 580.7, m 214-217° (others found m 228-230°).** The Mg complex can be prepared from protoporphyrin dimethyl ester (see below) and Mg(ClO<sub>4</sub>)<sub>2</sub> in boiling pyridine under N<sub>2</sub> for 3-4 hours until the band at 630nm (free porphyrin) is absent. Filter, wash the insolubles with Et<sub>2</sub>O until the filtrate is colourless. Evaporate the solvent *in vacuo* at 50-60° to a very small volume, then add excess peroxide free Et<sub>2</sub>O in a separating funnel, shake with H<sub>2</sub>O (2x), evaporate the Et<sub>2</sub>O, and remove the H<sub>2</sub>O and pyridine by evaporating (azeotropically) with \*C<sub>6</sub>H<sub>6</sub> (4x). Purify the residue by dissolving the Mg complex in as little hot \*C<sub>6</sub>H<sub>6</sub> (50ml for 800mg) as possible, and add cold petroleum ether (b 30-60°), leave at room temperature until crystallisation begins, then further in a refrigerator to give twinned prisms, **m** >330°. Its IR has  $\nu_{\max}$  at 3080(w CH=CH<sub>2</sub>), 1610, 1698, 1740 (s CO<sub>2</sub>Me) cm<sup>-1</sup>, but no NH. Its UV (Et<sub>2</sub>O) has  $\lambda_{\max}$  ( $\epsilon$ ) at 588nm (19,000), 550nm (19,300), 417nm (Soret, 252,000). [Ramsey *Biochemical Preparations* **3** 39 1953, Fuhrhop & Graniek *Biochemical Preparations* **13** 55 1971.]

**6-Mercaptopurine-9- $\beta$ -D-ribofuranoside** [574-25-4] **M 284.3, m 208-210°(dec), 210-211°(dec), 220-223°(dec), 222-224°(dec),  $[\alpha]_{\text{D}}^{25}$  -73° (c 1, 0.1N NaOH), pK<sup>25</sup> 7.56.** Recrystallise the riboside from H<sub>2</sub>O or EtOH. It has UV with  $\lambda_{\max}$  (H<sub>2</sub>O) at 322nm (pH 1), 320 nm (pH 6.7) and 310nm (pH 13). [IR: Johnson et al. *J Am Chem Soc* **80** 699 1958, UV: Fox et al. *J Am Chem Soc* **80** 1669 1958, *Beilstein* **26** III/IV 2100.]

**R(-)-Methadone (Levomethadone, 6-dimethylamino-4,4-diphenylheptan-3-one)** [125-58-6] **M 309.4, m 98-100°,  $[\alpha]_{\text{D}}^{20}$  -32° (c 1.8, EtOH), see below for pKa.** This pharmacologically active (against narcotic addiction)

enantiomer was obtained by optical resolution (using D-tartaric acid) of the racemate, and was purified by precipitation of the hydrochloride from aqueous solution at pH >6, dried and recrystallised from propan-2-ol. The *R*-hydrochloride [5967-73-7], when recrystallised from propan-2-ol, has *m* 245-246°,  $[\alpha]_{\text{D}}^{20}$  -169° (c 2, EtOH). The *S*-(+)-enantiomer [5653-8-5] also recrystallises from propan-2-ol and has recorded *m* of 100-101°,  $[\alpha]_{\text{D}}^{25}$  +26° (c 1.2, H<sub>2</sub>O). The *S*-hydrochloride [15284-15-8], when crystallised from propan-2-ol, has *m* 243-244°,  $[\alpha]_{\text{D}}^{20}$  +169° (c 2, EtOH). [Larsen et al. *J Am Chem Soc* **70** 4194 1948, Brode & Hill *J Org Chem* **13** 191 1948, Schultz et al. *J Am Chem Soc* **69** 2454 1947, Easton et al. *J Am Chem Soc* **69** 2941 1947, Winter & Flataker *J Pharmacol Exp Ther* **98** 305 1950, Beilstein **14** III 278, **14** III/IV 300.]

(±)-Methadone hydrochloride (6-dimethylamino-4,4-diphenylheptan-3-one HCl) [1095-90-5] *M* 345.9, *m* 241-242°,  $\text{pK}_{\text{I}}^{25}$  8.94,  $\text{pK}_{\text{2}}^{20}$  10.12 (free base). The salt (see above) crystallises from EtOH, or EtOH/Et<sub>2</sub>O. [See methadone references in the above entry.]

Methoxantin coenzyme (PQQ, pyrrolo quinoline quinone, 2,7,9-tricarboxy-1*H*-pyrrolo-[2,3-*f*]-quinoline-4,5-dione, 4,5-dihydro-4,5-dioxo-1*H*-pyrrolo[2,3-*f*]quinoline-2,7,9-tri-carboxylic acid) [72909-34-3] *M* 330.2, *m* 220°(dec). Efflorescent yellow-orange needles of PQQ are formed on recrystallising from H<sub>2</sub>O by addition of Me<sub>2</sub>CO, or better from a supersaturated aqueous solution, as it forms an acetone adduct. [Forrest et al. *Nature* **280** 843 1979.] It has also been purified by passage through a C-18 reverse phase silica cartridge or a silanised silica gel column in aqueous solution whereby methoxantin remains behind as a red-orange band at the origin. This band is collected and washed thoroughly with dilute aqueous HCl (pH 2) and is then eluted with MeOH/H<sub>2</sub>O (7:3) and evaporated *in vacuo* to give the coenzyme as a red solid. It has also been purified by dissolving it in aqueous 0.5M K<sub>2</sub>CO<sub>3</sub> and acidified to pH 2.5 whereby PQQ precipitates as a deep red solid which is collected and dried *in vacuo*. Methoxantin elutes at 3.55 retention volumes from a C18 μBondapak column using H<sub>2</sub>O/MeOH (95:5) + 0.1% AcOH pH 4.5. It has UV with  $\lambda_{\text{max}}$  at 247 and 330nm (shoulder at 270nm) in H<sub>2</sub>O and with  $\lambda_{\text{max}}$  at 250 and 340nm in H<sub>2</sub>O at pH 2.5. With excitation at  $\lambda_{\text{ex}}$  365nm it has a  $\lambda_{\text{em}}$  at 483nm. The <sup>13</sup>C NMR has  $\delta$ : 113.86, 122.76, 125.97, 127.71, 130.68, 137.60, 144.63, 146.41, 147.62, 161.25, 165.48, 166.45, 173.30 and 180.00.

When a solution in 10% aqueous MeCO is adjusted to pH 9 with aqueous NH<sub>3</sub> and kept at 25° for 30 minutes, the *acetone adduct* is formed, which has UV with  $\lambda_{\text{max}}$  at 250, 317 and 360nm (H<sub>2</sub>O, pH 5.5), and with  $\lambda_{\text{ex}}$  at 360nm it fluoresces with  $\lambda_{\text{em}}$  at 465nm; and the <sup>13</sup>C NMR [(CD<sub>3</sub>)<sub>2</sub>SO, TMS] has  $\delta$ : 29.77, 51.06, 74.82, 111.96, 120.75, 121.13, 125.59, 126.88, 135.21, 139.19, 144.92, 161.01, 161.47, 165.17, 168.61, 190.16 and 207.03. It also forms a *methanol adduct*.

When it is reacted with Me<sub>2</sub>SO<sub>4</sub>/K<sub>2</sub>CO<sub>3</sub> in dry Me<sub>2</sub>NCHO at 80° for 4 hours, it forms the *trimethyl ester* which has *m* 265-267°(dec) [260-263°(dec) also reported] after recrystallisation from hot MeCN it forms orange crystals with UV that has  $\lambda_{\text{max}}$  at 252 and 344nm (H<sub>2</sub>O) and 251, 321 and 373nm (in MeOH; MeOH adduct?). [Duine et al. *Eur J Biochem* **108** 187 1980, Duine et al. *Adv Enzymology* **59** 169 1987, Corey & Tramontano *J Am Chem Soc* **103** 5599 1981, Gainor & Weinreb *J Org Chem* **46** 4319 1981, Hendrickson & de Vries *J Org Chem* **17** 1148 1982, McKenzie et al. *J Chem Soc, Chem Commun* 1372 1983.]

Methyl benzylpenicillinate [653-89-4] *M* 348.3, *m* 97°,  $[\alpha]_{\text{D}}^{20}$  +328° (c 1, MeOH),  $[\alpha]_{\text{D}}^{20}$  +286° (c 1, CHCl<sub>3</sub>). Crystallise the ester from CCl<sub>4</sub> or EtOAc/hexane. [Sheehan et al. *J Am Chem Soc* **84** 2983 1962, Review: Chain et al. *Antibiotics* (Oxford University Press) **2** 1949.]

1-Methylimidazolidine-2,4-dione (1*N*-methylhydantoin, MH) [616-04-6] *M* 114.1, *m* 158°, 155-156°, 157-159°, and 184-185° (dimorphic ?),  $\text{pK}^{25}$  9.20. It is purified by dissolving it in the minimum volume of H<sub>2</sub>O, extracting with EtOAc, the extract is dried and evaporated to dryness. The residue, in EtOAc is subjected to column chromatography (silica gel with EtOAc as eluent), or silica gel TLC [MeOH/CHCl<sub>3</sub> (1:9)] and recrystallised from EtOH (elongated plates). It is identified by <sup>1</sup>H NMR and MS. The IR has  $\nu_{\text{max}}$  at 1712 (2CO) and 1761 (4CO) cm<sup>-1</sup>. The 3-*acetate* has *m* 134-135°. [Miller & Robson *J Chem Soc* 1910 1938, West *J Biol Chem* **34** 187 1918, Ienaga et al. *Tetrahedron Lett* **28** 4587 1987, Ienaga et al. *J Chem Soc, Perkin Trans I* 1153 1989, Beilstein **24** H 44.] Data kindly supplied by Dr Kazu Ienaga, Nippon Zoki Pharmaceutical Co Ltd, Osaka, Japan.

(±)-5-Methylimidazoline-2,4-dione (5-methylhydantoin) [616-03-5] *M* 114.1, *m* 145°, 148-152°, 149-151°,

**150°**, **pK<sup>25</sup> 9.07**. Purify it by recrystallisation from H<sub>2</sub>O, and dry it *in vacuo*. [Ienaga et al. *Tetrahedron Lett* **28** 4587 1987, Ienaga et al. *J Chem Soc, Perkin Trans I* 1153 1989, IR: Burland & Christian *Can J Chem* **35** 444 1957, *Beilstein* **24** H 279, **24** I 305, **24** II 155, **24** III/IV 1083.] Data supplied by Dr Kazu Ienaga (see above).

**5-Methyltetrahydrofolic acid disodium salt (prefolic A)** [68792-52-9] **M 503.4**, **pK<sub>1</sub> 2.4** (N10 protonation), **pK<sub>2</sub> 2.7** (pyrimidine N1 protonation), **pK<sub>3</sub> 3.5** ( $\alpha$ -CO<sub>2</sub>H), **pK<sub>4</sub> 4.9** ( $\gamma$ -CO<sub>2</sub>H), **pK<sub>5</sub> 5.6** (N5-Me), **pK<sub>6</sub> 8.5** (3NHCO acidic). First check the purity by measuring the UV at pH 7.0 (use phosphate buffer), and it should have  $\lambda_{\max}$  at 290nm and  $\lambda_{\min}$  at 245nm with a ratio of A<sub>290</sub>/A<sub>250</sub> of 3.7. This ratio goes down to 1.3 as oxidation to the dihydro derivative occurs. The latter can be reduced back to the tetrahydro compound by reaction with 2-mercaptoethanol at room temperature. If oxidation has occurred, then the compound should be chromatographed on DEAE-cellulose (~0.9 milliequiv/g, in AcO<sup>-</sup> form) in (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub> (1.5 M) and washed with 1M NH<sub>4</sub>OAc containing 0.01M mercaptoethanol till free from UV absorption and then washed with 0.01M mercaptoethanol. All is done in a nitrogen atmosphere. The reduced folate is then eluted with a gradient between 0.01M mercaptoethanol and 1M NH<sub>4</sub>OAc containing 0.01M mercaptoethanol, and the fractions with absorption at 290nm are collected. These are evaporated under reduced pressure at 25°, and traces of NH<sub>4</sub>OAc and H<sub>2</sub>O are removed at high vacuum/25° (~24-48 hours). The residue is dissolved in the minimum volume of 0.01M mercaptoethanol, and an equivalent of NaOH is added to convert the acid to the diNa salt and evaporated to dryness at 25°/high vacuum. The pure product has  $\lambda_{\max}$  290nm ( $\epsilon$  32,000) in pH 7.0 buffer. [Sakami *Biochemical Preparations* **10** 103 1963.]

**Mevalonic acid lactone** [674-26-0] **M 130.2**, **m 28°**, **b 145-150°/5mm**. Purify the lactone *via* the *dibenzylethylenediammonium salt* (**m** 124-125°) [Hofmann et al. *J Am Chem Soc* **79** 2316 1957], or by chromatography on paper or on a Dowex-1 (formate) column. [Bloch et al. *J Biol Chem* **234** 2595 1959.] Store it as the *N,N'*-dibenzylethylenediamine (DBED) salt, or as the lactone in a sealed container at 0°. [*Beilstein* **18/1** V 19.]

**Mevalonic acid 5-phosphate** [1189-94-2] **M 228.1**, **pK<sub>Est(1)</sub> ~1.5** (PO<sub>4</sub>H<sub>2</sub>), **pK<sub>Est(2)</sub> ~4.4** (CO<sub>2</sub>H), **pK<sub>Est(3)</sub> ~6.31** (PO<sub>4</sub>H<sup>-</sup>). Purify the acid by conversion to the *tricyclohexylammonium salt* (**m** 154-156°) by treatment with cyclohexylamine. Recrystallise the salt from water/acetone at -15°. *Alternatively*, the phosphate is chromatographed through an ion-exchange resin or paper (Whatman No 1) in a system of isobutyric acid/ammonia/water (66:3:30; R<sub>F</sub> 0.42). Store it as the cyclohexylammonium salt.

**Mevalonic acid 5-pyrophosphate** [1492-08-6] **M 258.1**, **pK<sub>Est(1)</sub> ~2**, **pK<sub>Est(2)</sub> ~2**, **pK<sub>Est(3)</sub> ~3.95** (PO<sub>4</sub>), **pK<sub>Est(4)</sub> 4.4** (CO<sub>2</sub>H), **pK<sub>Est(5)</sub> ~6.26** (PO<sub>4</sub>). Purify the pyrophosphate by ion-exchange chromatography on Dowex-1 formate [Bloch et al. *J Biol Chem* **234** 2595 1959], DEAE-cellulose [Skilletar and Kekwick, *Anal Biochem* **20** 171 1967], or by paper chromatography [Rogers et al. *Biochem J* **99** 381 1966]. Likely impurities are ATP and mevalonic acid phosphate. Store it as a dry powder or in a slightly alkaline (pH 7-9) solution at -20°.

**Mithramycin A (Aureolic acid, Plicamycin)** [18378-89-7] **M 1085.2**, **m 180-183°**,  $[\alpha]_D^{20}$  -51° (c 0.3, EtOH), **pK<sub>Est</sub> ~9.2**. Purify mithramycin A by crystallisation from CHCl<sub>3</sub>. It is soluble in MeOH, EtOH, Me<sub>2</sub>CO, EtOAc, Me<sub>2</sub>SO and H<sub>2</sub>O, and moderately soluble in CHCl<sub>3</sub>, but is slightly soluble in \*C<sub>6</sub>H<sub>6</sub> and Et<sub>2</sub>O. It is a fluorescent antitumour agent used in flow cytometry. [Thiem & Meyer *Tetrahedron* **37** 551 1981, NMR: Yu et al. *Nature* **218** 193 1968, *Beilstein* **17/1** V 672.]

**Mitomycin C** [50-07-7] **M 334.4**, **m >360°**, **pK<sub>Est(2)</sub> ~8.0**. Mitomycin C forms blue-violet crystals from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether. It is soluble in Me<sub>2</sub>CO, MeOH and H<sub>2</sub>O, moderately soluble in \*C<sub>6</sub>H<sub>6</sub>, CCl<sub>4</sub> and Et<sub>2</sub>O but insoluble in petroleum ether. It has UV with  $\lambda_{\max}$  at 216, 360 and a weak peak at 560nm in MeOH. [Stevens et al. *J Med Chem* **8** 1 1965, Shirahata & Hirayama *J Am Chem Soc* **105** 7199 1983, *Beilstein* **25** III/IV 516.]

**Muramic acid [R-2(2-amino-2-deoxy-D-glucose-3-yloxy)-propionic acid]** [1114-41-6] **M 251.2**, **m 145-150°(dec)**, **152-154°(dec)**, **155°(dec)**,  $[\alpha]_D^{25}$  +109° (c 2, H<sub>2</sub>O), +165.0° (extrapolated to 0 time) → +123° [after 3 hours (c 3, H<sub>2</sub>O)], **pK<sub>Est(1)</sub> ~3.8** (CO<sub>2</sub>), **pK<sub>Est(2)</sub> ~7.7** (NH<sub>2</sub>). Muramic acid crystallises from H<sub>2</sub>O or aqueous EtOH as the *monohydrate* which loses H<sub>2</sub>O at 80° *in vacuo* over P<sub>2</sub>O<sub>5</sub>. It sometimes contains some NaCl. It has

been purified by dissolving 2.2g in MeOH (75ml), filtering from some insoluble material, concentrating to ~10ml and refrigerating. The colourless crystals are washed with absolute MeOH. This process does not remove NaCl; to do so, the product is recrystallised from an equal weight of H<sub>2</sub>O to give a low recovery yield of very pure acid (0.12g). On paper chromatography 0.26µg give one ninhydrin positive spot after development with 75% phenol (R<sub>F</sub> 0.51) or with *sec*-BuOH/HCO<sub>2</sub>H/H<sub>2</sub>O (7:1:2) (R<sub>F</sub> 0.30). [Matsushima & Park *Biochemical Preparations* **10** 109 1963, *J Org Chem* **27** 3581 1962.] The acid has also been purified by dissolving 990mg in 50% aqueous EtOH (2ml), cooling, collecting the colourless needles on a sintered glass funnel and drying over P<sub>2</sub>O<sub>5</sub> at 80°/0.1mm to give the anhydrous acid. [Lambert & Zilliken *Chem Ber* **93** 2915 1960.] Alternatively the acid is dissolved in a small volume of H<sub>2</sub>O, neutralised to pH 7 with ion-exchange resin beads (IR4B in OH<sup>-</sup> form), filtered, evaporated and dried. The residue is recrystallised from 90% EtOH (v/v) and dried as above for 24 hours. [Strange & Kent *Biochem J* **71** 333 1959.] The *N*-acetyl derivative (**NAMA**, ***R*-2-(acetylamino)-3-*O*-(1-carboxyethyl)-2-deoxy-D-glucose, *R*-2(2-acetylamino-2-deoxy-D-glucose-3-yloxy)-propionic acid**) [10597-89-4], M 292.3, has *m* ~125° (dec) and [α]<sub>D</sub><sup>20</sup> +41.2° after 24 hours (c 1.5, H<sub>2</sub>O), pK<sub>Est</sub> ~ 3.6. [Watanabe & Saito *J Bacteriol* **144** 428 1980, *Beilstein* **4** IV 2029.]

**Muscimol (pantherine, 5-aminoethyl-3[2*h*]-isoxazolone)** [2763-96-4] M 114.1, *m* 170-172°(dec), 172-174°(dec), 172-175°, 175°, 176-178°(dec), pK<sub>Est(1)</sub> ~ 6 (acidic, ring 2-NH), pK<sub>Est(2)</sub> ~ 8 (CH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub>). Recrystallise muscimol from MeOH/tetrahydrofuran or EtOH and sublime it at 110-140° (bath) at 10<sup>-4</sup> mm to give a yellow spot with ninhydrin which slowly turns purple [NMR: Bowden et al. *J Chem Soc (C)* 172 1968]. It can also be purified by dissolving in the minimum volume of hot H<sub>2</sub>O and adding EtOH dropwise until cloudy, cool, and colourless crystals separate; its IR has ν<sub>□□□</sub> at 3445w, 3000-2560w br, 2156w, 1635s and 1475s cm<sup>-1</sup>. [NMR: Jager & Frey *Justus Liebigs Ann Chem* 817 1982.] Alternatively, it has been purified by two successive chromatographic treatments on Dowex-1 x 8, with the first elution with 2M AcOH and a second with a linear gradient between 0—2M AcOH, evaporating the desired fractions and recrystallising the residue from MeOH. [McCarry & Savard *Tetrahedron Lett* **22** 5153 1981, Nakamura *Chem Pharm Bull Jpn* **19** 46 1971.]

**Mycophenolic acid (6-[1,3-dihydro-7-hydroxy-5-methoxy-4-methyl-1-oxoisobenzofuran-6-yl]-4-methylhex-4-enoic acid)** [24280-93-1] M 320.3, *m* 141°, 141-143°, pK<sub>Est(1)</sub> ~ 2.5 (CO<sub>2</sub>H), pK<sub>Est(2)</sub> ~ 9.5 (phenolic OH). Purify the acid by dissolving it in the minimum volume of EtOAc, applying onto a silica gel column (0.05-0.2 mesh) and eluting with a mixture of EtOAc/CHCl<sub>3</sub>/AcOH (45:55:1) followed by recrystallisation from heptane/EtOAc, from aqueous EtOH or from hot H<sub>2</sub>O and drying *in vacuo*. It is a weak dibasic acid, moderately soluble in Et<sub>2</sub>O, CHCl<sub>3</sub> and hot H<sub>2</sub>O but weakly soluble in \*C<sub>6</sub>H<sub>6</sub> and toluene. [Birch & Wright *Aust J Chem* **22** 2635 1969, Canonica et al. *J Chem Soc, Perkin Trans 1* 2639 1972, Birkinshaw et al. *Biochem J* **50** 630 1952, *Beilstein* **18** II 393, **18** III/IV 6513.]

**Myricetin (Cannabiscetin, 3,3',4',5,5',7-hexahydroxyflavone)** [529-44-2] M 318.2, *m* >300°, 357°(dec) (polyphenolic pK<sub>Est</sub> ~ 8-11). Recrystallise myricetin from aqueous EtOH (*m* 357° dec, as *monohydrate*) or Me<sub>2</sub>CO (*m* 350° dec, with one mol of Me<sub>2</sub>CO) as yellow crystals. It is almost insoluble in CHCl<sub>3</sub> and AcOH. The *hexaacetate* has *m* 213°. [Hergert *J Org Chem* **21** 534 1956, Spada & Cameroni *Gazzetta* **86** 965, 975 1956, Kalf & Robinson *J Chem Soc* **127** 181 1925, *Beilstein* **18/5** V 670.]

**Nalidixic acid (1-ethyl-7-methyl-1,8-naphthyridin-4-one-3-carboxylic acid)** [389-08-2] M 232.3, *m* 226.8-230.2°, 228-230°, 229-230°, pK<sup>25</sup> 6.0. Nalidixic acid crystallises from H<sub>2</sub>O or EtOH as a pale buff powder. It is soluble at 23° in CHCl<sub>3</sub> (3.5%), toluene (0.16%), MeOH (0.13%), EtOH (0.09%), H<sub>2</sub>O (0.01%) and Et<sub>2</sub>O (0.01%). It inhibits nucleic acid and protein synthesis in yeast. [Leshner et al. *J Med & Pharm Chem* **5** 1063 1962.]

**Naloxone hydrochloride hydrate (Narcan, 1-*N*-propenyl-7,8-dihydro-14-hydroxymorphinan-6-one hydrochloride)** [51481-60-8] M 399.9, *m* 200-205°, [α]<sub>D</sub><sup>20</sup> -164° (c 2.5, H<sub>2</sub>O), pK<sub>Est(1)</sub> ~ 6 (*N*-propenyl), pK<sub>Est(2)</sub> ~ 9.6 (phenolic OH). This opiate antagonist has been recrystallised from EtOH/Et<sub>2</sub>O or H<sub>2</sub>O. It is soluble in H<sub>2</sub>O (5%) and EtOH but insoluble in Et<sub>2</sub>O. The *free base* has *m* 184° (177-178° also reported) after recrystallisation from EtOAc, and [α]<sub>D</sub><sup>20</sup> -194.5° (c 0.93, CHCl<sub>3</sub>). [Olofson et al. *Tetrahedron Lett* 1567 1977,

Gold et al. *Med Res Rev* 2 211 1982.]

**Naltrexone hydrochloride dihydrate (1-N-cyclopropylmethyl-7,8-dihydro-14-hydroxy-morphinan-6-one hydrochloride)** [16676-29-2] **M 413.9, m 274-276°**,  $[\alpha]_D^{20} -173^\circ$  (c 1, H<sub>2</sub>O), **pK<sub>Est(1)</sub> ~6 (N-cyclopropylmethyl), pK<sub>Est(2)</sub> ~9.6 (phenolic OH)**. This narcotic antagonist has been purified by recrystallisation from MeOH and dried in air. The *free base* has **m 168-170°** after recrystallisation from Me<sub>2</sub>CO. [Cone et al. *J Pharm Sci* 64 618 1975, Gold et al. *Med Res Rev* 2 211 1982.]

**1-Naphthyl phosphate disodium salt** [2183-17-7] **M 268.1, pK<sub>1</sub><sup>26</sup> 0.97, pK<sub>2</sub><sup>26</sup> 5.85 (for free acid)**. Purify the salt through an acid ion-exchange column (in H<sup>+</sup> form) to give the *free acid* [1136-89-6], **M 224.2**, which is obtained by freeze drying and recrystallising from Me<sub>2</sub>CO/\*C<sub>6</sub>H<sub>6</sub>, or by adding 2.5 volumes of hot CHCl<sub>3</sub> (or 20 parts of boiling \*C<sub>6</sub>H<sub>6</sub>) to a hot solution of 1 part acid and 1.2 parts Me<sub>2</sub>CO and cooling (**m 155-157°**, 157-158°). The acid is dissolved in the minimum volume of H<sub>2</sub>O to which 2 equivalents of NaOH are added and then freeze dried, or by adding the equivalent amount of MeONa in MeOH to a solution of the acid in MeOH and collecting the Na salt, washing with cold MeOH, then Et<sub>2</sub>O, and drying in a vacuum. [Friedman & Seligman *J Am Chem Soc* 72 624 1950, Chanley & Fegeason *J Am Chem Soc* 77 4002 1955.] The *monosodium salt* [1136-89-6] is similarly prepared but by using 1 equivalent of NaOH. The phosphate group is hydrolysed at pH 1.1-5.85 at 70°. [Beilstein 6 IV 4226.] It is a substrate for alkaline phosphatase [Gomori *Methods Enzymol* 4 381 1957, 128 212 1968], and prostatic phosphatase [Babson *Clin Chem* 30 1418 1984]. [See p 519 in "Miscellaneous As, B, P, Si, S, Se and Te Compounds", Chapter 4, *Beilstein* 6 IV 4226.]

**2-Naphthyl phosphate monosodium salt** [14463-68-4] **M 246.2, m 177-178°, 296° (sintering at 228°), pK<sub>1</sub><sup>26</sup> 1.28, pK<sub>2</sub><sup>26</sup> 5.53, pK<sub>3</sub><sup>26</sup> 6.57 (for free acid)**. The *free acid* [41845-15-2] is purified as for the preceding 1-isomer and has **m 176-177°** (also 177-178°) after several recrystallisations by adding 2.5 volumes of hot CHCl<sub>3</sub> to a hot solution of 1 part of acid in 1.3 volumes of Me<sub>2</sub>CO as for the 1-isomer above. It is neutralised with one equivalent of NaOH and freeze dried or prepared as the 1-isomer above. Its solubility in H<sub>2</sub>O is ~5%. It also forms a 0.5 Na.1 H<sub>2</sub>O salt which has **m 203-205°** (244° also reported). [Friedman & Seligman *J Am Chem Soc* 72 624 1950, Chanley & Fegeason *J Am Chem Soc* 77 4002 1955, See p 520 in "Miscellaneous As, B, P, Si, S, Se and Te Compounds", Chapter 4, *Beilstein* 6 IV 4285.]

**D(+)-Neopterin** [2009-64-5] **M 253.2, m >300°(dec)**,  $[\alpha]_{546}^{20} +64.5^\circ$  (c 0.14, 0.1M HCl),  $[\alpha]_D^{25} +50.1^\circ$  (c 0.3, 0.1N HCl), **pK<sub>1</sub> 2.23 (basic), pK<sub>2</sub> 7.89 (acidic)**. Purification is as for biopterin. Also purify it on a Dowex-1 x 8 (formate form) column and elute with 0.03M ammonium formate buffer pH 8.0 then pH 7.2. The fluorescent neopterin fraction is evaporated under reduced pressure, leaving neopterin and ammonium formate (the latter sublimes out at high vacuum) behind. Stir the residue for 24 hours with EtOH, collect the solid and recrystallise it from H<sub>2</sub>O. [Viscontini et al. *Helv Chim Acta* 53 1202 1970, cf Wachter et al. Eds *Neopterin* W de Gruyter, Berlin 1992, ISBN 9783110117905, *Beilstein* 26 IV 4038.]

**β-Nicotinamide adenine dinucleotide (diphosphopyridine nucleotide, NAD, DPN)** [53-84-9] **M 663.4,  $[\alpha]_D^{23} -34.8^\circ$  (c 1, H<sub>2</sub>O), pK<sub>1</sub> 2.2 (PO<sub>3</sub>H), pK<sub>2</sub> 4.0 (adenine NH<sub>2</sub>), pK<sub>3</sub> 6.1 (PO<sub>4</sub>)**. NAD is purified by paper chromatography or better on a Dowex-1 ion-exchange resin. The column is prepared by washing with 3M HCl until free of material absorbing at 260nm, then with water, 2M sodium formate until free of chloride ions and, finally, with water. NAD, as a 0.2% solution in water, is adjusted with NaOH to pH 8, and adsorbed onto the column, washed with water, and eluted with 0.1M formic acid. Fractions with strong absorption at 360nm are combined, acidified to pH 2.0 with 2M HCl, and cold acetone (*ca* 5L/g of NAD) is added slowly and with constant agitation. It is left overnight in the cold, then the precipitate is collected in a centrifuge, washed with pure acetone and dried under vacuum over CaCl<sub>2</sub> and paraffin wax shavings [Kornberg *Methods Enzymol* 3 876 1957]. It has been purified by anion-exchange chromatography [Dalziel & Dickinson *Biochemical Preparations* 11 84 1966.] The purity is checked by reduction to NADH (with EtOH and yeast alcohol dehydrogenase) which has  $\epsilon_{340\text{nm}} 6220 \text{ M}^{-1}\text{cm}^{-1}$ . [Todd et al. *J Chem Soc* 3727, 3733 1957.] [pKa, Lamborg et al. *J Biol Chem* 231 685 1958.] The *free acid* crystallises from aqueous Me<sub>2</sub>CO with 3H<sub>2</sub>O and has **m 140-142°**. It is stable in cold neutral aqueous solutions in a desiccator (CaCl<sub>2</sub>) at 25°, but decomposes at strong acid and alkaline pH. Its purity is checked by reduction with yeast alcohol dehydrogenase and EtOH to NADH and noting the OD at 340nm. Pure NADH (see below) has  $\epsilon_{340} 6.2 \times 10^4 \text{ M}^{-1}\text{cm}^{-1}$ , i.e. 0.1μmole of NADH in 3ml and in a 1cm path

length cell has an OD at 340nm of 0.207. [*Beilstein* 26 IV 3644.]

**$\beta$ -Nicotinamide adenine dinucleotide reduced di-Na salt trihydrate (reduced diphosphopyridine nucleotide sodium salt, NADH)** [606-68-8] **M 763.5, pKa as for NAD.** This coenzyme is available in high purity, and it is advisable to buy a fresh preparation rather than to purify an old sample as purification will invariably lead to a more impure sample contaminated with the oxidised form (NAD). It has UV with  $\lambda_{\max}$  at 340nm ( $\epsilon$  6,200 M<sup>-1</sup>cm<sup>-1</sup>) at which wavelength the oxidised form NAD has no absorption. At 340nm a 0.161mM solution in a 1cm (pathlength) cell has an absorbance of 1.0 unit. The purity is best checked by the ratio  $A_{280\text{nm}}/A_{340\text{nm}} \sim 2.1$ , a value which increases as oxidation proceeds. The dry powder is stable indefinitely at -20°. Solutions in aqueous buffers at pH  $\sim 7$  are stable for extended periods at -20° and for at least 8 hours at 0°, but are oxidised more rapidly at 4° in a cold room (e.g. almost completely oxidised overnight at 4°). [UV: Drabkin *J Biol Chem* 175 563 1945, Fluorescence: Boyer & Thorell *Acta Chem Scand* 10 447 1956, Redox: Rodkey *J Biol Chem* 234 188 1959, Schlenk in *The Enzymes* 2 250, 268 1951, Kaplan in *The Enzymes* 3 105, 112 1960.] Deuterated NADH, i.e. NADD, has been purified through the anion exchange resin AG-1 x 8 (100-200 mesh, formate form) and through a Bio-Gel P-2 column. [Viola et al. *Anal Biochem* 96 334 1979, *Beilstein* 26 III/IV 3639.]

**$\beta$ -Nicotinamide adenine dinucleotide phosphate (NADP, TPN)** [53-59-8] **M 743.4, pK<sub>1</sub> 1.1 (PO<sub>4</sub>H<sub>2</sub>), pK<sub>2</sub> 4.0 (adenine NH<sub>2</sub>), pK<sub>3</sub> 6.1 (PO<sub>4</sub> ).** Purify it by anion-exchange chromatography in much the same way as for NAD [Dalziel & Dickinson *Biochem J* 95 311 1965, *Biochemical Preparations* 11 87 1966]. Finally it is purified by dissolving in H<sub>2</sub>O and precipitating with 4 volumes of Me<sub>2</sub>CO and dried *in vacuo* over P<sub>2</sub>O<sub>5</sub>. It is unchanged by storing *in vacuo* at 2°. [Hughes et al. *J Chem Soc* 3733 1957, Schuster & Kaplan *J Biol Chem* 215 183 1955.] Deuterated NADPH, i.e. NADPD, has been purified through the anion exchange resin AG-1 x 8 (100-200 mesh, formate form) and through a Bio-Gel P-2 column.  $\lambda_{\min}$  259nm ( $\epsilon$  18.000) at pH 7.0. [Viola et al. *Anal Biochem* 96 334 1979, *Beilstein* 26 IV 3669, 3672.]

**$\beta$ -Nicotinamide adenine dinucleotide phosphate reduced tetrasodium salt (reduced diphosphopyridine nucleotide phosphate sodium salt, NADPH)** [2646-71-1] **M 833.4, pKa as for NADP.** Purification is mostly similar to that of NADH above. [*Beilstein* 26 III/IV 3671.]

**$\beta$ -Nicotinamide mononucleotide (NMN)** [1094-61-7] **M 334.2,  $[\alpha]_D^{23} -38.3^\circ$  (c 1, H<sub>2</sub>O), pK<sub>Est</sub>  $\sim 6.1$  (PO<sub>4</sub> ).** Purify NMN by passage through a column of Dowex-1 (Cl<sup>-</sup> form) and washing with H<sub>2</sub>O until no absorbance is observed at 260 nm. The tubes containing NMN are pooled, adjusted to pH 5.5-6 and evaporated *in vacuo* to a small volume. This is adjusted to pH 3 with dilute HNO<sub>3</sub> in an ice-bath and treated with 20 volumes of Me<sub>2</sub>CO at 0-5°. The heavy white precipitate is collected by centrifugation at 0°. It is best stored wet and frozen or it can be dried to give a gummy residue. It has  $\lambda_{\max}$  at 266nm ( $\epsilon$  4,600) and  $\lambda_{\min}$  at 249nm ( $\epsilon$  3600) at pH 7.0 (i.e. no absorption at 340nm). It can be estimated by reaction with CN<sup>-</sup> or hydrosulfite which form the 4-adducts (equivalent to NADH) which have UV with  $\lambda_{\max}$  at 340nm ( $\epsilon$  6,200). Thus after reaction, an OD<sub>340</sub> of "one" is obtained from a 0.1612mM solution in a 1cm path cuvette. [Plaut & Plaut *Biochemical Preparations* 5 56 1957, Maplan & Stolzenbach *Methods Enzymol* 3 899 1957, Kaplan et al. *J Am Chem Soc* 77 815 1955, *Beilstein* 22/2 V 168.]

**(-)-Nicotine ((2S)-1-methyl-2[3-pyridyl]-pyrrolidine)** [54-11-5] **M 162.2, b 123-125°/17mm, 243-248°/760mm (partial dec),  $d_4^{20}$  1.097,  $n_D^{20}$  1.5280,  $[\alpha]_D^{20} -169^\circ$  (c 1, Me<sub>2</sub>CO), pK<sub>1</sub><sup>15</sup> 6.16 (pyridine N<sup>+</sup>), pK<sub>2</sub><sup>15</sup> 10.96 (pyrrolidine N<sup>+</sup>).** (-)-Nicotine is a very pale yellow *hygroscopic* oil with a characteristic odour (tobacco extract) which turns brown in air on exposure to light. It is purified by fractional distillation under reduced pressure in an inert atmosphere. A freshly distilled sample should be stored in dark sealed containers under N<sub>2</sub>. It is a strong base; a 0.05 M aqueous solution has a pH of 10.2. It is very soluble in organic solvents. It is soluble in H<sub>2</sub>O and readily forms salts. [UV: Parvis *J Chem Soc* 97 1035 1910, Dobbie & Fox *J Chem Soc* 103 1194 1913.] The *hydrochlorides* (mono- and di-) form deliquescent crystals soluble in H<sub>2</sub>O and EtOH but insoluble in Et<sub>2</sub>O. It has also been purified *via* the ZnCl<sub>2</sub> double salt. [Ratz *Monatsh Chem* 26 1241 1905, Biosynthesis: Nakan & Hitchinson *J Org Chem* 43 3922 1978.] The *picrate* has **m** 218° (from EtOH). [*Beilstein* 23/6 V 64.] **POISONOUS.**

**(±)-Nicotine** [22083-74-5] **M 162.2, b 113-115°/10mm, 242.3°/atm,  $d_4^{20}$  1.082 (pKa see above).** It is purified

by fractional distillation. Its solubility in EtOH is ~5%. The *picrate* forms yellow needles from hot H<sub>2</sub>O and has **m** 219°. The *methiodide* has **m** 219° (from MeOH). [Craig *J Am Chem Soc* **55** 2854 1933, Nakane & Hutchinson *J Org Chem* **43** 3922 1978, *Beilstein* **23/6** V 64.] **POISONOUS.**

**Nonactin** [6833-84-7] **M** 737.0, **m** 147-148°,  $[\alpha]_{\text{D}}^{20}$  0° (±2°) (c 1.2, CHCl<sub>3</sub>). This macrotetrolide antibiotic crystallises from MeOH as colourless needles and is dried at 90°/20 hours/high vacuum. [Cordaz et al. *Helv Chim Acta* **38** 1445 1955, crystal structure: Dobler *Helv Chim Acta* **55** 1371 1972, Gambos et al. *Tetrahedron Lett* 3391 1975, *Beilstein* **19/12** V 751.]

**L-Noradrenaline (Adrenor, R-2-amino-1-[3,4-dihydroxyphenyl]ethan-1-ol, L-norepi-nephriene)** [51-41-2, 69815-49-2, 636-88-4 (*bitartrate salt*)] **M** 169.2, **m** 216.5-218°(dec), ~220-230°(dec),  $[\alpha]_{\text{D}}^{20}$  -45° (c 5, N HCl),  $[\alpha]_{\text{D}}^{25}$  -37.3° (c 5, 1 equivalent aqueous HCl), **pK**<sub>1</sub><sup>25</sup> 5.58 (phenolic OH), **pK**<sub>2</sub><sup>25</sup> 8.90 (phenolic OH), **pK**<sub>3</sub><sup>25</sup> 9.78 (NH<sub>2</sub>). Recrystallise adrenor from EtOH and store it in the dark under N<sub>2</sub>. [pKa, Lewis *Brit J Pharmacol Chemother* **9** 488 1954, UV: Bergström et al. *Acta Physiol Scand* **20** 101 1950, Fluorescence: Bowman et al. *Science NY* **122** 32 1955, Tullar *J Am Chem Soc* **70** 2067 1948.] The *L-tartrate salt monohydrate* has **m** 102-104.5°,  $[\alpha]_{\text{D}}^{25}$  -11° (c 1.6, H<sub>2</sub>O), after recrystallisation from H<sub>2</sub>O or EtOH. [*Beilstein* **13** III 2382.]

**L-Noradrenaline hydrochloride (Arterenol)** [329-56-6] **M** 205.6, **m** 145.2-146.4°, ~150°(dec),  $[\alpha]_{\text{D}}^{25}$  -40° (c 6, H<sub>2</sub>O), **pKa** see above. Recrystallise arterenol from isoPrOH and store it in the dark as it is oxidised under light (see preceding entry). [Tullar *J Am Chem Soc* **70** 2067 1948, *Beilstein* **13** III 2382.]

**1R,2S(-)-Norephedrine [L(-)-erythro-1R,2S-2-amino-1-phenyl-1-propanol]** [492-41-1] **M** 151.2, **m** ~49-53°, 50-52°,  $[\alpha]_{\text{D}}^{25}$  -14.6° (c 3.4, EtOH),  $[\alpha]_{\text{D}}^{20}$  -41° (c 7, M HCl), **pK**<sup>25</sup> 8.92. It crystallises in plates from H<sub>2</sub>O or Et<sub>2</sub>O/petroleum ether.

The *1R,2S(-) hydrochloride* **M** 187.7, has **m** 174-175°, crystallises from iso-PrOH, and has  $[\alpha]_{\text{D}}^{20}$  -33.0° (c 5, H<sub>2</sub>O). [cf Atkins & Cramer *J Am Chem Soc* **52** 4349 1930, *Beilstein* **13** II 370, **13** III 1717, **13** IV 1875.]

**1S,2R(+)-Norephedrine [D(+)-erythro-1R,2S-2-amino-1-phenyl-1-propanol]** [37577-28-9] **M** 151.2, **m** ~49-53°, 50-52°,  $[\alpha]_{\text{D}}^{27}$  +14.8° (c 4, EtOH),  $[\alpha]_{\text{D}}^{20}$  +40° (c 7, M HCl), **pK**<sup>25</sup> 8.92. Purify it by recrystallisation from H<sub>2</sub>O (plates), and it is soluble in Et<sub>2</sub>O. Lewis [*Brit J Pharmacol Chemother* **9** 488 1954] obtained **pK**<sub>2</sub><sup>20</sup> 9.44 (H<sub>2</sub>O).

The *hydrochloride* [1S,2R(+)] 40626-28-7] **M** 187.7, has **m** 172-175° (plates from EtOH) and  $[\alpha]_{\text{D}}^{20}$  +33.4° (c 6, H<sub>2</sub>O), and the *sulfate* (plates from H<sub>2</sub>O) has **m** 285-286° and  $[\alpha]_{\text{D}}^{20}$  +31.5° (H<sub>2</sub>O). [cf Atkins & Cramer *J Am Chem Soc* **52** 4349 1930, *Beilstein* **13** II 371, **13** III 1717, **13** IV 1875.]

**DL-(±)-Norephedrine hydrochloride [Propadrin, (±)-erythro-1RS,2SR-2-amino-1-phenyl-1-propanol hydrochloride]** [154-41-6] **m** 194-196° (also 194° was reported), **pK** 8.20. It crystallises in plates from absolute EtOH or isoPrOH, and the (±)-*oxalate* has **m** 245°(dec) (plates from H<sub>2</sub>O). The *2,4-dinitrobenzoate* has **m** 86-88° (from EtOH) and the *3,5-dinitrobenzoate* has **m** 78-79° (from EtOH). It is a mixed anti-sympathomimetic and used as a nasal anticongestant. [cf Atkins & Cramer *J Am Chem Soc* **52** 4349 1930, Krantz & Hartung *J Am Pharm Assoc* **19** 461 1930, Fischer & Plein *J Am Pharm Assoc* **44** 313 1955, *Beilstein* **13** I 252, **13** II 371, **13** III 1717, **13** IV 1875.]

**1R,2R(-)-Norpseudoephedrine [L(-)-threo-1R,2R-2-amino-1-phenyl-1-propanol]** [37577-07-4] **M** 151.2, **m** ~50°, 50-52°, 77°, 77.5-78°,  $[\alpha]_{\text{D}}^{20}$  -34° (c 3.5, EtOH), **pK**<sup>25</sup> 8.92. Purify (-)-nor-ψ-ephedrine by recrystallisation from H<sub>2</sub>O, MeOH, EtOH, Et<sub>2</sub>O/petroleum ether or \*C<sub>6</sub>H<sub>6</sub> (plates). The *mandelate salt* has **m** 163.5° (from EtOH/Et<sub>2</sub>O) and  $[\alpha]_{\text{D}}^{32}$  -41.3° (c 0.8, H<sub>2</sub>O) [Jarowski & Hartung *J Org Chem* **8** 564 566 567 1943]. The *hydrochloride* is purified by dissolving 1.44g in 96% EtOH (5ml), adding Et<sub>2</sub>O (16ml) and cooling; it has **m** 178-179° (**m**180-181° is also reported) and  $[\alpha]_{\text{D}}^{30}$  -42.9° (c 1.8, H<sub>2</sub>O) [Fles & Markovac-Prpic *Croat Chem Acta* **29** 186 1957]. [*Beilstein* **13** I 252, **13** II 370, **13** III 1716, **13** IV 1874.]

**(+)-(1S,2S)-Norpseudoephedrine [(+)-Cathine, D(+)-threo-(1S,2S)-2-amino-1-phenyl-1-propan-1-ol]** [492-39-7] **M** 151.2, **m** 77.5-78°,  $[\alpha]_{\text{D}}^{25}$  +34.0° (c 4, EtOH). Recrystallise it from \*C<sub>6</sub>H<sub>6</sub> and store away from CO<sub>2</sub> as it is a strong base readily forming a *hydrochloride* [1S,2S(+)] 2153-98-2] **M** 187.7, **m** 180-181° (prisms from

EtOH), and is soluble in H<sub>2</sub>O with  $[\alpha]_D^{20} +42.5^\circ$ . It is a stimulant and is an anorexic substance. [Nagaai & Kanao *Justus Liebigs Ann Chem* **470** 157 1929, Sicher & Pankova *Col Czech Chem Commun* **20** 1409 1955, *Beilstein* **13** I 252, **13** II 370, **13** III 1716, **13** IV 1874.]

**Novobiocin** (7-[O<sup>3</sup>-carbamoyl-5-O<sup>4</sup>-dimethyl-β-L-lyso-6-desoxyhexahydropyranosyloxy]-4-hydroxy-3[4-hydroxy-3-{3-methylbut-2-enyl}-benzylamino]-8-methylcoumarin) [303-81-1] **M 612.6**, two forms **m 152-156°** and **m 172-174°, 174-178°, λ<sub>max</sub> at 330nm (acidic EtOH), 305nm (alkaline EtOH), [α]<sub>D</sub><sup>25</sup> -63° (c 1, EtOH), pK<sub>1</sub> 4.03 (4.2), pK<sub>2</sub> 9.16**. Crystallise novobiocin from EtOH and store it in the dark. It has also been recrystallised from Me<sub>2</sub>CO/H<sub>2</sub>O. [Hoeksema et al. *J Am Chem Soc* **77** 6710 1955, Kaczka et al. *J Am Chem Soc* **77** 9404 1955.]

The **sodium salt** [1476-53-5] **M 634.6**, **m 210-215°, 215-220°(dec), 222-229°, [α]<sub>D</sub><sup>25</sup> -38° (c 1, H<sub>2</sub>O)** has been recrystallised from MeOH, then dried at 60°/0.5mm. [Sensi et al. *Anal Chem* **29** 1611 1957, Kaczka et al. *J Am Chem Soc* **78** 4126 1956, *Beilstein* **18/8** IV 8125.]

**Nucleotide thiophosphate analogues.** The preparation and purification of [<sup>3</sup>H]ATPγS, [<sup>3</sup>H]GTPγS, S<sup>6</sup>ITPγS (6-thioinosine), Cl<sup>6</sup>ITPγS (6-chloroinosine) and [<sup>3</sup>H]ATPγS were described, and general purification was achieved by chromatography of the nucleotide thiophosphates in the minimum volume of H<sub>2</sub>O placed onto a DEAE-Sephadex A25 column and eluted with a linear gradient of triethylammonium bicarbonate (0.1 to 0.6M for G and I nucleotides and 0.2 to 0.5M for A nucleotides). [Googy et al. *Biochim Biophys Acta* **276** 155 1972.]

**Nystatin dihydrate (Mycostatin, Fungicidin)** [1400-61-9] **M 962.1**, **m dec>160° (without melting by 250°), [α]<sub>D</sub><sup>25</sup> -7° (0.1N HCl in MeOH), -10° (AcOH), +12° (Me<sub>2</sub>NCHO), +21° (pyridine)**. Nystatin is a light yellow powder with the following solubilities at ~28°: MeOH (1.1%), ethylene glycol (0.9%), H<sub>2</sub>O (0.4%), CCl<sub>4</sub> (0.12%), EtOH (0.12%), CHCl<sub>3</sub> (0.05%) and \*C<sub>6</sub>H<sub>6</sub> (0.03%). It has been precipitated from MeOH solution by addition of H<sub>2</sub>O. Aqueous suspensions of this macrolide antifungal antibiotic are stable at 100°/10minutes at pH 7.0 but decompose rapidly at pH <2 and >9, and in the presence of light and O<sub>2</sub>. [Birch et al. *Tetrahedron Lett* 1491, 1485 1964, Weiss et al. *Antibiot Chemother* **7** 374 1957.] It may contain a mixture of components A<sub>1</sub>, A<sub>2</sub> and A<sub>3</sub>. [*Beilstein* **18** III/IV 7480.]

**Oxacillin sodium salt (5-methyl-3-phenyl-4-isoxazolympenicillin sodium salt)** [1173-88-2] **M 423.4**, **m 188°(dec), [α]<sub>D</sub><sup>20</sup> +29° (c 1, H<sub>2</sub>O), pK<sub>Est</sub> ~2.7**. This antibiotic, which is stable to penicillinase, is purified by recrystallisation from isoPrOH and dried *in vacuo*. Its solubility in H<sub>2</sub>O at 25° is 5%. [Doyle et al. *Nature* **192** 1183 1961, Review: Chain et al. *Antibiotics* (Oxford University Press) **2** 1949.]

**Oxiracetam (4-hydroxy-2-oxo-1-pyrrolidine acetamide)** [62613-82-5 unspecified, 68567-97-5 racemate] **M 158.2**, **m 165-168°**. This nootropic (Alzheimer) drug is purified by recrystallisation from MeOH or aqueous Me<sub>2</sub>CO and dried *in vacuo*. [NMR, IR: Pifferi & Pinza *Farmaco Ed Sci* **32** 602 1977, Banfi et al. *Farmaco Ed Sci* **39** 16 1984, Gouilaev & Senning *Brain Research Rev* **19** 180 1994.]

**R-(+)-oxiracetam** [68252-28-8] **m 135-136°, [α]<sub>D</sub><sup>20</sup> +36° (c 1, H<sub>2</sub>O)**, and **S-(-)-oxiracetam** [88929-35-5] **m 135-136°, [α]<sub>D</sub><sup>20</sup> +36.5° (c 1, H<sub>2</sub>O)**, have been recrystallised from aqueous Me<sub>2</sub>CO. [Gouilaev & Senning *Brain Research Rev* **19** 180 1994.]

**Palmitoyl coenzyme A** [1763-10-6] **M 1005.9**. Possible impurities are palmitic acid, S-palmitoyl thioglycolic acid and S-palmitoyl glutathione. These are removed by placing *ca* 200mg in a centrifuge tube and extracting with Me<sub>2</sub>CO (20ml), followed by two successive extractions with Et<sub>2</sub>O (15ml) to remove S-palmitoyl thioglycolic acid and palmitic acid. The residue is dissolved in H<sub>2</sub>O (4 x 4 ml), adjusted to pH 5 and centrifuged to remove insoluble S-palmitoyl glutathione and other insoluble impurities. To the clear supernatant is added 5% HClO<sub>4</sub> (6ml) whereby S-palmitoyl CoA precipitates. The precipitate is washed with 0.8% HClO<sub>4</sub> (10ml) and finally with Me<sub>2</sub>CO (3x 5ml) and dried *in vacuo*. It is stable for at least one year in dry form at 0° in a desiccator (dark). Solutions are stable for several months at -15°. Its solubility in H<sub>2</sub>O is 4%. The adenine content is used as the basis of purity with λ<sub>max</sub> at 260 and 232nm (ε 6.4 x 10<sup>6</sup> and 9.4 x 10<sup>6</sup> cm<sup>2</sup>/mol, respectively). Higher

absorption at 232nm would indicate other thio ester impurities, e.g. S-palmitoyl glutathione, which absorb highly at this wavelength. Also the phosphate content should be determined, and acid phosphate can be titrated potentiometrically. [Seubert *Biochemical Preparations* 7 80 1960, Srer et al. *Biochim Biophys Acta* 33 31 1959, Kornberg & Pricer *J Biol Chem* 204 329, 345 1953, Beilstein 26 III/IV 3665.]

**3-Palmitoyl-*sn*-glycerol** (*R*-glycerol-1-palmitate, *L*- $\beta$ -palmitin) [32899-41-5] **M 330.5,  $d_{27.3}^{20}$  0.9014, m 66.5° ( $\alpha$ -form), 74° ( $\beta'$ -form) and 77° ( $\beta$ -form).** The stable  $\beta$ -form is obtained by crystallisation from EtOH or Skellysolve B, and recrystallisation from Et<sub>2</sub>O provides the  $\beta'$ -form. The  $\alpha$ -form is obtained on cooling the melt. [Malkin & el Sharbagy *J Chem Soc* 1631 1936, Chapman *J Chem Soc* 58 1956, Luton & Jackson *J Am Chem Soc* 70 2446 1948, Beilstein 2 III 966.]

**D-Panthenol** (Provitamin B, *R*-2,4-dihydroxy-3,3-dimethylbutyric acid 3-hydroxy-propylamide) [81-13-0] **M 205.3, b 118-120°/0.02mm,  $d_{20}^{20}$  1.2,  $n_D^{20}$  1.4935,  $[\alpha]_D^{20} +30^\circ$  (c 5, H<sub>2</sub>O).** Purify D-panthenol by distillation *in vacuo*. It is a slightly *hygroscopic* viscous oil and is soluble in H<sub>2</sub>O and organic solvents. It is hydrolysed by alkali and strong acid. [Rabin *J Am Pharm Assoc (Sci Ed)* 37 502 1948, Bonati & Pitre *Farmaco Ed Scient* 14 43 1959, Beilstein 4 IV 1652.]

***R*-(+)-Pantothenic acid sodium salt** (*N*-[2,4-dihydroxy-3,3-dimethylbutyryl]  $\beta$ -alanine Na salt) [867-81-2] **M 241.2,  $[\alpha]_D^{25} +27.1^\circ$  (c 2, H<sub>2</sub>O),  $pK_{25}^{25}$  4.4 (for free acid).** Crystallise the salt from absolute EtOH. It is very *hygroscopic* (keep in sealed ampoules). The *free acid* is a viscous *hygroscopic* oil with  $[\alpha]_D^{25} +37.5^\circ$  (c 5, H<sub>2</sub>O), easily destroyed by acids and bases. See next entry. [Beilstein 4 IV 2569.]

***R*-(+)-Pantothenic acid Ca salt** [(D(+)) 137-08-6, 63409-48-3] **M 476.5, m 195-196°, 200-201°,  $[\alpha]_D^{20} +28.2^\circ$  (c 5, H<sub>2</sub>O).** The salt crystallises as needles from MeOH, EtOH or isoPrOH (with 0.5mol of isoPrOH). It is moderately *hygroscopic*. The *S*-benzylisothiuronium salt has **m 151-152° (149° when crystallised from Me<sub>2</sub>CO).** [Kagan et al. *J Am Chem Soc* 79 3545 1957, Wilson et al. *J Am Chem Soc* 76 5177 1954, Stiller & Wiley *J Am Chem Soc* 63 1239 1941, Beilstein 4 IV 2569.]

**Paromomycin sulfate** {amminosidin, *O*-2,6-diamino-2,6-dideoxy- $\beta$ -*L*-idopyranosyl-(1 $\rightarrow$ 3)-*O*- $\beta$ -*D*-ribofuranosyl(1 $\rightarrow$ 5)-*O*-[2-amino-2-deoxy- $\alpha$ -*D*-glucopyranosyl-(1 $\rightarrow$ 4)]-2-deoxystrepta amine sulfate} [1263-89-4] **M 713.7, amorphous,  $[\alpha]_D^{25} +50^\circ$  (c 1.5, H<sub>2</sub>O pH6),  $pK_{Est} \sim 9$ .** Purify it by dissolving it in H<sub>2</sub>O (0.5g/10ml) and adding excess EtOH, filter or collect and wash with EtOH, then Et<sub>2</sub>O by centrifugation, and dry it *in vacuo*. An aqueous solution is stable at 37° for a week but longer at 0-5°. The *free base* [7542-37-2] is a white amorphous powder which should be stored under N<sub>2</sub> because it is strongly basic and can absorb CO<sub>2</sub> from the atmosphere. It is soluble in MeOH (less soluble in EtOH) and has  $[\alpha]_D^{25} +65^\circ$  (c 1.5, MeOH). It is an antimicrobial against Gram +ve and Gram -ve bacteria and is antiamoebic. It inhibits initiation and peptide elongation during protein synthesis. [Haskell et al. *J Am Chem Soc* 81, 3480, 3482 1959, Hichens & Rinehart *J Am Chem Soc* 85 1547 1963, Beilstein 18 III/IV 7534.]

**D(-)-Penicillamine** (*R*-3-mercapto-*D*-valine, 3,3-dimethyl-*D*-cysteine, from natural penicillin) [52-67-5] **M 149.2, m 202-206°, 214-217°,  $[\alpha]_D^{21} -63^\circ$  (c 1, N NaOH or pyridine).** The melting point of D(-)-penicillamine depends on the rate of heating (**m 202-206°** is obtained by starting at 195° and heating at 2°/minute). It is soluble in H<sub>2</sub>O and alcohols but insoluble in Et<sub>2</sub>O, CHCl<sub>3</sub>, CCl<sub>4</sub> and hydrocarbon solvents. Purify it by dissolving it in MeOH and adding Et<sub>2</sub>O slowly. Dry it *in vacuo* and store it under N<sub>2</sub>. [Weight et al. *Angew Chem, Int Ed (English)* 14 330 1975, Cornforth in *The Chemistry of Penicillin* (Clarke, Johnson and Robinson eds) Princeton Univ Press, 455 1949, Review: Chain et al. *Antibiotics* (Oxford University Press) 2 1949, Polymorphism: Vidler *J Pharm Pharmacol* 28 663 1976]. The *D*-*S*-benzyl derivative has **m 197-198° (from H<sub>2</sub>O),  $[\alpha]_D^{17} -20^\circ$  (c 1, N NaOH), -70° (N HCl).** [Beilstein 4 IV 3228.]

***L*-(-)-Penicillamine** [1113-41-3] **M 149.2, m 190-194°, 202-206°, 214-217°,  $[\alpha]_D^{21} +63^\circ$  (c 1, N NaOH or pyridine).** Same as the preceding entry for its enantiomer. [Beilstein 4 IV 3228.]  
***D*-Penicillamine disulfide hydrate** (*S,S'*-di-*D*-penicillamine) hydrate) [20902-45-8] **M 296.4 + aqueous, m 203-204°(dec), 204-205°(dec),  $[\alpha]_D^{23} +27^\circ$  (c 1.5, N HCl), -82° (c 0.8, N NaOH),  $pK_{Est(1)} \sim 2.4$  (CO<sub>2</sub>),  $pK_{Est(2)}$**

~ 10.7 (NH<sub>2</sub>). Purify it by recrystallisation from EtOH or aqueous EtOH. [Crooks in *The Chemistry of Penicillin* (Clarke, Johnson and Robinson Eds) Princeton University Press, 469 1949. Use as a thiol reagent for proteins: Garel *Eur J Biochem* **123** 513 1982, Süss *Justus Liebigs Ann Chem* **561** 31 1948, *Beilstein* **4** IV 3231.]

**Penicillic acid (5-hydroxy-5-isopropenyl-4-methoxy-2(5H)furanone)** [90-65-3] **M 158.2, m 58-64°, 64-65° (monohydrate, acid), 83-84°, 87° (anhydrous, lactone), pK<sup>25</sup> 5.9.** The lactone (furanone, *anhydrous*) hydrolyses to the acid (3-methoxy-4-oxo-hexa-2,5-dienoic acid, *hydrate*). It crystallises from water as the *monohydrate* (acid), or from petroleum ether as the *anhydrous* (lactone) compound. The free acid is in equilibrium with the lactone. Its UV has  $\lambda_{\max}$  at 221nm ( $\epsilon$  12,500) in 0.02M KOH; 228nm ( $\epsilon$  11,500) in 0.02M HCl [Raphael *J Chem Soc* 1508 1948]. [*Beilstein* **3** II 519, **3** III 1467.] It is a possible **antineoplastic**.

**3-sn-Phosphatidylethanolamine (L- $\alpha$ -cephalin, from Soya bean)** [39382-08-6] **M<sub>r</sub> ~600-800, amorphous, pK<sub>Est(1)</sub>~5.8 (PO<sub>4</sub><sup>-</sup>), pK<sub>Est(2)</sub>~10.5 (NH<sub>2</sub>).** Purify the cephalin by dissolving it in EtOH, adding Pb(OAc)<sub>2</sub>·3H<sub>2</sub>O (30g in 100ml H<sub>2</sub>O) until excess Pb<sup>2+</sup> is present. Filter off the solid. Pass CO<sub>2</sub> gas through the solution until precipitation of PbCO<sub>3</sub> ceases. Filter the solid off and evaporate (while bubbling CO<sub>2</sub>) under vacuum. An equal volume of H<sub>2</sub>O is added to the residual oil and extracted with hexane. The hexane extract is washed with H<sub>2</sub>O until the aqueous phase is free from Pb [test with dithizone (2 mg in 100 ml CCl<sub>4</sub>; Feigl *Spot Tests* Vol I, Elsevier p. 10 1954)]. The hexane is dried (Na<sub>2</sub>SO<sub>4</sub>), filtered and evaporated to give a yellow waxy solid which should be dried to constant weight *in vacuo*. It is practically insoluble in H<sub>2</sub>O and Me<sub>2</sub>CO, but freely soluble in CHCl<sub>3</sub> (5%) and Et<sub>2</sub>O, and slightly soluble in EtOH. [Schofield & Dutton *Biochemical Preparations* **5** 5 1957.]

**O-Phosphocolamine (2-aminoethyl dihydrogen phosphate)** [1071-23-4] **M 141.1, m 237-240°, 242.3°, 234.5-244.5°, 244-245°(capillary), pK<sub>1</sub><sup>20</sup> <1.5 (PO<sub>4</sub>H<sub>2</sub>), pK<sub>2</sub><sup>20</sup> 5.77 (PO<sub>4</sub>H<sup>-</sup>), pK<sub>3</sub><sup>20</sup> 10.26 (NH<sup>+</sup>).** Purification by recrystallisation from aqueous EtOH gives a *hydrate* (**m** 140-141°). Its solubility in H<sub>2</sub>O is 17% and 0.003% in MeOH or EtOH at 22°. [Fölisch & Österberg *J Biol Chem* **234** 2298 1959, Baer & Stauber *Can J Chem* **34** 434 1956, Christensen *J Biol Chem* **135** 399 1940.] It is a potent inhibitor of ornithine decarboxylase [Gilad & Gilad *Biochem Biophys Res Commun* **122** 277 1984]. [*Beilstein* **4** IV 1415.]

**Phosphoenolpyruvic acid monopotassium salt (KPEP)** [4265-07-0] **M 206.1, pK<sub>1</sub><sup>25</sup> 3.4 (CO<sub>2</sub>), pK<sub>2</sub><sup>25</sup> 6.35 (PO<sub>4</sub>H<sup>-</sup>) (for free acid).** KPEP is purified *via* the monocyclohexylamine salt (see next entry). The salt (534mg) in H<sub>2</sub>O (10ml) is added to Dowex 50Wx4 H<sup>+</sup> form (200-400 mesh, 2ml, H<sub>2</sub>O washed) and stirred gently for 30 minutes and filtered. The resin is washed with H<sub>2</sub>O (6ml), and the combined solutions are adjusted to pH 7.4 with 3N KOH (~1.4ml) and the volume adjusted to 18.4ml with H<sub>2</sub>O to give a solution of 0.1M KPEP which can be lyophilised to a pure powder and is very good for enzyme work. It has been recrystallised from MeOH/Et<sub>2</sub>O. [Clark & Kirby *Biochemical Preparations* **11** 103 1966, Wold & Ballou *J Biol Chem* **227** 301 1957, Cherbuliez & Rabinowitz *Helv Chim Acta* **39** 1461 1956, *Beilstein* **3** IV 977.]

**The triNa salt** [5541-93-5] **M 360.0,** is purified as follows: the salt (1g) is dissolved in MeOH (40ml) and dry Et<sub>2</sub>O is added in excess. The white crystals are collected and dried over P<sub>2</sub>O<sub>5</sub> at 20°. [Cramer & Voges *Chem Ber* **92** 952 1959, *Beilstein* **3** IV 977.]

**Phosphoenolpyruvic acid tris(cyclohexylamine) salt** [35556-70-8] **M 465.6, m 155-180°(dec).** Recrystallise it from aqueous Me<sub>2</sub>CO and dry it in a vacuum. At 4° it is stable for >2 years and has IR at 1721cm<sup>-1</sup> (C=O). [Wold & Ballou *J Biol Chem* **227** 301 1957, Clark & Kirby *Biochemical Preparations* **11** 103 1966 for the monocyclohexylamine salt, *Beilstein* **12** IV 8.]

**D-3-Phosphoglyceric acid disodium salt (D-glycerate 3-phosphate di-Na salt)** [80731-10-8] **M 230.0, [ $\alpha$ ]<sub>D</sub><sup>25</sup> +7.7° (c 5, H<sub>2</sub>O), -735° (in aqueous NH<sub>4</sub><sup>+</sup> molybdate), pK<sub>Est(1)</sub>~1.0 (PO<sub>4</sub>H<sub>2</sub>), pK<sub>Est(2)</sub>~6.66 (PO<sub>4</sub>H<sup>-</sup>) (for free acid).** It is best purified by conversion to the **Ba salt** by precipitation with BaCl<sub>2</sub>, which is recrystallised three times before conversion to the sodium salt. The **Ba salt** (9.5g) is shaken with 200ml of a 1:1 slurry of Dowex 50 (Na<sup>+</sup> form) for 2 hours. The mixture is filtered, and the resin is washed with H<sub>2</sub>O (2 x 25ml). The combined filtrates (150ml) are adjusted to pH 7.0 and concentrated *in vacuo* to 30-40ml and filtered if not clear. Absolute EtOH is added to make 100ml, and then *n*-hexane is added whereby a white solid and/or a second phase separates. When set aside at room temperature, complete precipitation of the Na salt as a solid occurs. The salt is

removed by centrifugation, washed with Me<sub>2</sub>CO, dried in air then in an oven at 55° to give a stable powder (4.5g). It did not lose weight when dried further over P<sub>2</sub>O<sub>5</sub> at 78°/8 hours. The high rotation in the presence of (NH<sub>4</sub>)<sub>6</sub>Mo<sub>7</sub>O<sub>24</sub> is not very sensitive to the concentration of molybdate or pH as it did not alter appreciably in 1/3 volume between 2.5 to 25% (w/v) of molybdate or at pH values ranging between 4 and 7. [Cowgill *Biochim Biophys Acta* **16** 613 1955, Embdan et al. *Hoppe Seyler's Z Physiol Chem* **230** 20 1934, *Beilstein* **3** IV 1051.]

**Phospholipids.** For the removal of ionic contaminants from raw zwitterionic phospholipids, most lipids are purified twice by mixed-bed ionic exchange resins (Amberlite AB-2) of methanolic solutions. (About 1g of lipid in 10ml of MeOH). With both runs the first 1ml of the eluate is discarded. The main fraction of the solution is evaporated at 40°C under dry N<sub>2</sub> and recrystallised three times from *n*-pentane. The resulting white powder is dried for about 4 hours at 50° under reduced pressure and stored at 3°. Some samples are purified by mixed-bed ion exchange of aqueous suspensions of the crystal/liquid crystal phase. [Kaatze et al. *J Phys Chem* **89** 2565 1985.]

**O-Phospho-L-serine** [407-41-0] **M 185.1, m 175-176°**, [α]<sub>D</sub><sup>20</sup> +4.3° (c 3.2, H<sub>2</sub>O), +16.2° (c 3.2, 2N HCl), pK<sub>i</sub><sup>25</sup> <1 (PO<sub>4</sub>H<sub>2</sub>), pK<sub>2</sub><sup>25</sup> 2.08 (CO<sub>2</sub>H), pK<sub>3</sub><sup>25</sup> 5.65 (PO<sub>4</sub>H<sup>-</sup>), pK<sub>4</sub><sup>25</sup> 9.74 (NH<sub>3</sub><sup>+</sup>). Recrystallise the phosphoserine by dissolving 10g in H<sub>2</sub>O (150ml) at 25°, stirring for up to 20 minutes. Undissolved material is filtered off (Büchner), and 95% EtOH (85ml) is added dropwise during 4 minutes, and set aside at 25° for 3 hours, then at 3° overnight. The crystals are washed with 95% EtOH (100ml), then dry Et<sub>2</sub>O (50ml), and dried in a vacuum (yield 6.5g). A further quantity (1.5g) can be obtained by keeping the mother liquors and washings at -10° for 1 week. The *DL*-isomer has **m** 167-170°(dec) after recrystallisation from H<sub>2</sub>O/EtOH or MeOH. [Neuhaus & Korke *Biochemical Preparations* **6** 75 1958, Neuhaus & Byrne *J Biol Chem* **234** 113 1959, IR: Fölsch & Mellander *Acta Chem Scand* **11** 1232 1957, *Beilstein* **4** IV 3120.]

**O-Phospho-L-threonine (L-threonine-O-phosphate)** [1114-81-4] **M 199.1, m 194°(dec)**, [α]<sub>D</sub><sup>24</sup> -7.4° (c 2.8, H<sub>2</sub>O) (pKa as above). Dissolve the phosphate in the minimum volume of H<sub>2</sub>O, add charcoal, stir for a few minutes, filter and apply onto a column of Dowex 50W (H<sup>+</sup> form), then elute with 2N HCl. Evaporate the eluates under reduced pressure whereby the desired fraction produces crystals of the phosphate which can be recrystallised from H<sub>2</sub>O/MeOH mixtures and the crystals are then dried *in vacuo* over P<sub>2</sub>O<sub>5</sub> at ~80°. [de Verdier *Acta Chem Scand* **7** 196 1953, *Beilstein* **4** IV 3175.]

**O-Phospho-L-tyrosine (L-tyrosine-O-phosphate)** [21820-51-9] **M 261.2, m 225°, 227°, 253°**, [α]<sub>D</sub><sup>20</sup> -5.5° (c 1, H<sub>2</sub>O), -9.2° (c 1, 2N HCl), pK<sub>Est(1)</sub> ~1.6 (PO<sub>4</sub>H<sub>2</sub>), pK<sub>Est(2)</sub> ~2.02 (CO<sub>2</sub>H), pK<sub>Est(3)</sub> ~5.65 (PO<sub>4</sub>H<sup>-</sup>), pK<sub>Est(4)</sub> 9.2 (NH<sub>3</sub><sup>+</sup>). Purify it by recrystallisation from H<sub>2</sub>O or H<sub>2</sub>O/EtOH. [Levene & Schormüller *J Biol Chem* **100** 583 1933, Posternak & Graff *Helv Chim Acta* **28** 1258 1945, *Beilstein* **14** III 1510.]

**Phytol (d-2E-3,7R,11R,15-tetramethylhexadec-2-en-1-ol)** [150-86-7, 7541-49-93] **M 296.5, b 145°/0.03mm, 150-151°/0.06mm, 202-204°/10mm, d<sub>4</sub><sup>25</sup> 0.8497, n<sub>D</sub><sup>25</sup> 1.437, [α]<sub>D</sub><sup>22</sup> +0.06° (neat)**. It is commercially available as a mixture of isomers. Phytol is purified by distillation under high vacuum. It is almost insoluble in H<sub>2</sub>O but soluble in most organic solvents. It has UV with λ<sub>max</sub> at 212nm (log ε 3.04) in EtOH and IR has ν<sub>max</sub> at 3300 and 1670cm<sup>-1</sup>. [Demole & Lederer *Bull Soc Chim Fr* 1128 1958, Burrell *J Chem Soc (C)* 2144 1966, Bader *Helv Chim Acta* **34** 1632 1951, *Beilstein* **1** IV 2208.]

**Piracetamide (2-oxo-1-pyrrolidine acetamide)** [7491-74-9] **M 142.2, m 151.5-152.5°**. This typical nootropic (Alzheimer) drug modulates Na flux in AMPA receptors and is purified by recrystallisation from isoPrOH. [Gouilaev & Senning *Brain Research Rev* **19** 180 1994.]

**Podophylotoxin** [518-28-5] **M 414.4, m 181-181°, 183-184°, 188-189°**, [α]<sub>D</sub><sup>20</sup> -132° (c 1, CHCl<sub>3</sub>). The toxin recrystallises from \*C<sub>6</sub>H<sub>6</sub> (with 0.5C<sub>6</sub>H<sub>6</sub>), EtOH/\*C<sub>6</sub>H<sub>6</sub>, aqueous EtOH (with 1-1.5H<sub>2</sub>O, **m** 114-115°) and CH<sub>2</sub>Cl<sub>2</sub>/pentane. When dried at 100°/10mm it has **m** 183-184°. [UV: Stoll et al. *Helv Chim Acta* **37** 1747 1954, IR: Schecler et al. *J Org Chem* **21** 288 1956.] It is an inhibitor of microtubule assembly [Prasad et al. *Biochemistry* **25** 739 1986]. [*Beilstein* **19/10** V 666.]

**Polyethylene glycol** [25322-68-3] **M<sub>r</sub> various, from PEG ~200 to ~35,000**. PEG is available commercially as

a powder or as a solution in various degrees of polymerisation depending on the average molecular weight, e.g. PEG 400 and PEG 800 have average molecular weights of 400 and 800, respectively. They may be contaminated with aldehydes and peroxides. Solutions deteriorate in the presence of air due to the formation of these contaminants. Methods available for purification are as follows:

*Procedure A:* A 40% aqueous solution of PEG 400 (2L, average molecular weight 400) is de-aerated under vacuum and made 10mM in sodium thiosulfate. After standing for 1 hour at 25°, the solution is passed through a column (2.5x20cm) of mixed-bed R-208 resin which has a 5cm layer of Dowex 50-H<sup>+</sup> at the bottom of the column. The column was previously flushed with 30% aqueous MeOH, then thoroughly with H<sub>2</sub>O. A flow rate of 1ml/minute is maintained by adjusting the fluid head. The first 200ml are discarded, and the effluent is then collected at an increased flow rate. The concentration of PEG solution is checked by density measurement, and it is stored (preferably anaerobically) at 15°.

*Procedure B:* A solution of PEG 800 (500g in 805ml H<sub>2</sub>O) is made 1mM in H<sub>2</sub>SO<sub>4</sub> and stirred overnight at 25° with 10g of treated Dowex 50-H<sup>+</sup> (8% crosslinked, 20-50 mesh). The resin, after settling, is filtered off on a sintered glass funnel. The filtrate is treated at 25° with 1.5g of NaBH<sub>4</sub> (added over a period of 1 minute) in a beaker with tight but removable lid through which a propeller-type mechanical stirrer is inserted and continuously flushed with N<sub>2</sub>. After 15 minutes, 15g of fresh Dowex 50-H<sup>+</sup> are added, and the rate of stirring is adjusted to maintain the resin suspended. The addition of an equal quantity of Dowex 50-H<sup>+</sup> is repeated and the reaction times are 30 and 40 minutes. The pH of a 1 to 10 dilution of the reaction mixture should remain above pH 8 throughout. If it does not, more NaBH<sub>4</sub> is added or the addition of Dowex 50-H<sup>+</sup> is curtailed. (Some samples of PEG can be sufficiently acidic, at least after the hydrolysis treatment, to produce a pH that is too low for efficient reduction when the above ratio of NaBH<sub>4</sub> to Dowex 50-H<sup>+</sup> is used.) About 30 minutes after the last addition of NaBH<sub>4</sub>, small amounts of Dowex 50-H<sup>+</sup> (~0.2g) are added at 15 minute intervals until the pH of a 1 to 10 dilution of the solution is less than 8. After stirring for an additional 15 minutes the resin is allowed to settle, and the solution is transferred to a vacuum flask for brief de-gassing under a vacuum. The de-gassed solution is passed through a column of mixed-bed resin as in procedure A. The final PEG concentration would be about 40% w/v.

Assays for aldehydes by the purpural method and of peroxides are given in the reference below.

*Treatment of Dowex 50-H<sup>+</sup> (8% crosslinked, 20-50 mesh):* The Dowex (500g) is suspended in excess 2N NaOH, and 3ml of liquid Br<sub>2</sub> is stirred into the solution. After the Br<sub>2</sub> has dissolved, the treatment is repeated twice, and then the resin is washed with 1N NaOH on a sintered glass funnel until the filtrate is colourless. The resin is then converted to the acid form (with dilute HCl, H<sub>2</sub>SO<sub>4</sub> or AcOH as required) and washed thoroughly with H<sub>2</sub>O and sucked dry on the funnel. The treated resin can be converted to the Na salt and stored.

[Ray & Purathalingal *Anal Biochem* **146** 307 1985.]

**Porphobilinogen (5-amino-4-carboxymethyl-1H-pyrrole-3-propionic acid)** [487-90-1] **M 226.2, m 172-175°(dec), 175-180°(dec, darkening at 120-130°), pK<sub>1</sub> 3.70 (4-CH<sub>2</sub>CO<sub>2</sub>H), pK<sub>2</sub> 4.95 (3-CH<sub>2</sub>CH<sub>2</sub>CO<sub>2</sub>H), pK<sub>3</sub> 10.1 (NH<sup>+</sup>).** Porphobilinogen recrystallises as the *monohydrate* (pink crystals) from dilute NH<sub>4</sub>OAc solutions of pH 4, and is dried *in vacuo*. The *hydrochloride monohydrate* has **m 165-170°(dec)** (from dilute HCl). [Jackson & MacDonald *Can J Chem* **35** 715 1957, Westall *Nature* **170** 614 1952, Bogarad *J Am Chem Soc* **75** 3610 1953, Beilstein **22/14 V** 210.]

**Porphyrin a (from ox heart)** [5162-02-1] **M 799.0, m dec on heating.** It is purified on a cellulose powder column followed by extraction with 17% HCl and fractionated with HCl. [Morell et al. *Biochem J* **78** 793 1961.] It recrystallises from CHCl<sub>3</sub>/petroleum ether or Et<sub>2</sub>O/\*C<sub>6</sub>H<sub>6</sub> [detailed UV-VIS and NMR data: Caughey et al. *J Biol Chem* **250** 7602 1975, Lemberg *Adv Enzymol* **23** 265 1961].

**Prazosin hydrochloride (2[4-{(2-furoyl)piperazin-1-yl}4-amino-6,7-dimethoxyquinazoline hydrochloride)** [19237-84-4] **M 419.9, m 278-280°, 280-282°, pK<sub>25</sub> 6.5.** The salt is recrystallised by dissolving it in hot MeOH, adding a small volume of MeOH/HCl (dry MeOH saturated with dry HCl gas) followed by dry Et<sub>2</sub>O until crystallisation is complete. Dry it *in vacuo* over solid KOH till the odour of HCl is absent. It has been recrystallised from hot H<sub>2</sub>O, the crystals are washed with H<sub>2</sub>O, and the H<sub>2</sub>O is removed azeotropically with CH<sub>2</sub>Cl<sub>2</sub>, and dried in a vacuum. [NMR and IR: Honkanen et al. *J Heterocycl Chem* **17** 797 1980, cf Armarego & Reece *Aust J Chem* **34** 1561 1981.] It is an antihypertensive drug and is an α<sub>1</sub>-adrenergic antagonist [Brosman et al. *Proc Natl Acad Sci USA* **82** 5915 1985].

**Procaine hydrochloride** (Novocain, 2-diethylaminoethyl-4-aminobenzoate) [51-05-8] M 272.8, m 153-156°, 154-156°, 156°,  $\text{pK}_{\text{Est}(1)} \sim 2.52$  ( $\text{NH}_2^+$ )  $\text{pK}_2^{20}$  9.0 ( $\text{Et}_2\text{N}^+$ ). Novocain is recrystallised from aqueous EtOH. It is soluble at 25° in  $\text{H}_2\text{O}$  (86.3%), EtOH (2.6%) and  $\text{Me}_2\text{CO}$  (1%), it is slightly soluble in  $\text{CHCl}_3$ , but is almost insoluble in  $\text{Et}_2\text{O}$ . The anhydrous *free base* is recrystallised from ligroin or  $\text{Et}_2\text{O}$  and has m 61°. [Einhorn *Justus Liebigs Ann Chem* 371 125 1909, IR: Szymanski & Panzica *J Amer Pharm Assoc* 47 443 1958, *Beilstein* 14 IV 1138.]

**R-Propranalol hydrochloride** (*R*-1-isopropylamino-3-(1-naphthoxy)-2-propanol HCl) [13071-11-9] M 295.8, m 192°, 193-195°,  $[\alpha]_{\text{D}}^{20}$  -25° (c 1, EtOH),  $\text{pK}^{20}$  9.5 (for *free base*). The hydrochloride is recrystallised from *n*-PrOH or  $\text{Me}_2\text{CO}$ . It is soluble in  $\text{H}_2\text{O}$  and EtOH but is insoluble in  $\text{Et}_2\text{O}$ ,  $\text{C}_6\text{H}_6$  or EtOAc. The *racemate* has m 163-164°, and the *free base* recrystallises from cyclohexane with m 96°. [Howe & Shanks *Nature* 210 1336 1966.] The *S*-isomer (below) is the physiologically active isomer.

**S-Propranalol hydrochloride** (*S*-1-isopropylamino-3-(1-naphthoxy)-2-propanol HCl) [4199-10-4] M 295.8, m 192°, 193-195°,  $[\alpha]_{\text{D}}^{20}$  +25° (c 1, EtOH)  $\text{pK}^{20}$  9.5. See preceding entry for physical properties and purification. The (+)-salt is the active isomer which blocks isoprenaline tachycardia and is a  $\beta$ -adrenergic blocker. [Leclerc et al. *Trends Pharmacol Sci* 2 18 1981, Howe & Shanks *Nature* 210 1336 1966.]

**Protoporphyrin IX** (3,18-divinyl-2,7,13,17-tetramethylporphyrin-8,12-dipropionic acid, ooporphyrin) [553-12-8] M 562.7,  $\text{pK}_{\text{Est}} \sim 4.8$ . Protoporphyrin IX is purified by dissolving (4g) in 98-100%  $\text{HCOOH}$  (85ml), diluting with dry  $\text{Et}_2\text{O}$  (700ml) and keeping at 0° overnight. The precipitate is collected and washed with  $\text{Et}_2\text{O}$ , then  $\text{H}_2\text{O}$ , and dried in a vacuum at 50° over  $\text{P}_2\text{O}_5$ . It crystallises from aqueous pyridine and from  $\text{Et}_2\text{O}$  in monoclinic, brownish-yellow prisms. The UV has  $\lambda_{\text{max}}$  at values in 25% HCl are 557.2, 582.2 and 602.4nm. It is freely soluble in ethanolic HCl, AcOH,  $\text{CHCl}_3$ , and  $\text{Et}_2\text{O}$  containing AcOH. It forms sparingly soluble diNa and diK salts. [Ramsey *Biochemical Preparations* 3 39 1953, UV: Holden *Aust J. Exptl Biol and Med Sci* 15 412 1937, Garnick *J Biol Chem* 175 333 1948, IR: Falk & Willis *Aust J Sci Res [A]* 4 579 1951, *Beilstein* 26 IV 3042.]

**Protoporphyrin IX dimethylester** [5522-66-7] M 590.7, m 228-230°. The crude dimethyl ester (1g) in  $\text{CHCl}_3$  (200 ml) is mixed with petroleum ether (b 70-90°, 600ml), and any porphyrin (m > 260°) which is insoluble in this mixture is filtered off. The filtrate is passed through a column of  $\text{CaCO}_3$  [from  $\text{CaCO}_3$  (130g) which is kept overnight in a mixture of  $\text{CHCl}_3$ /petroleum ether (b 70-90°, 1:3), and the slurry is poured into a glass tube (2.5 x 26cm) to form the column]. After all the filtrate is applied, the column is eluted with a solution of  $\text{CHCl}_3$ /petroleum ether (b 70-90°, 1:3). All the coloured eluates are collected, evaporated at room temperature in a vacuum to give a residue (0.8g), m 208-211°. The residue (0.8g) in  $\text{CHCl}_3$  (66ml) is heated briefly to its boiling point, then boiling MeOH (198ml) is added immediately to it. The mixture is allowed to cool to room temperature, refrigerated for 2 days and the solid is filtered off. The solid is washed on the filter funnel with  $\text{CHCl}_3$ /MeOH (1:9, 50ml) and dried at 50°/vacuum (yield 0.62-0.66g). It can also be recrystallised by dissolving in as little hot dry  $\text{C}_6\text{H}_6$  as possible and left overnight at 20°, and has m 228-230°.

It has also been purified by dissolving (0.4g) in  $\text{CHCl}_3$  (33ml) by boiling for a few minutes, then diluting with boiling MeOH (100ml) and refrigerating for 2 days. The crystals are collected, washed with  $\text{CHCl}_3$ /MeOH (1:9) and dried at 50° in a vacuum (yield 0.3g). Its UV has  $\lambda_{\text{max}}$  at 631, 576, 541, 506 and 407nm in  $\text{CHCl}_3$  and 601, 556 and 406nm in 25% HCl. [Ramsey *Biochemical Preparations* 3 39 1953, *Beilstein* 26 III/IV 3052.]

**Pterin-6-carboxylic acid** (2-amino-4-oxo-3,4-dihydropteridine-6-carboxylic acid) [948-60-7] M 207.2, m >360°,  $\text{pK}_1^{20}$  1.43,  $\text{pK}_2^{20}$  2.88,  $\text{pK}_3^{20}$  7.72. The acid gives yellow crystals by repeated dissolution in aqueous NaOH and adding aqueous HCl. It has UV with  $\lambda_{\text{max}}$  at 235, 260 and 265nm ( $\epsilon$  11,000, 10,500 and 9,000) in 0.1N HCl and 263 and 365nm ( $\epsilon$  20,500 and 9000) in 0.1N NaOH. [UV: Pfeleiderer et al. *Justus Liebigs Ann Chem* 741 64 1970, Stockstad et al. *J Am Chem Soc* 70 5 1948, Fluorescence: Kavanagh & Goodwin *Arch Biochem* 20 315 1949, *Beilstein* 26 III/IV 4053.]

**Purine-9- $\beta$ -ribofuranoside** (Nebularin) [550-33-4] M 252.2, m 178-180°, 181-182°,  $[\alpha]_{\text{D}}^{25}$  -48.6° (c 1,  $\text{H}_2\text{O}$ ), -22° (c 0.8, 0.1N HCl) and -61° (c 0.8, 0.1N NaOH),  $\text{pK}^{25}$  2.05. Nebularin is recrystallised from butanone/Mebutanone/MeOH or EtOH and forms a MeOH photo-adduct. It is a strong inhibitor of adenosine

deaminase [EC 3.5.4.4]. [Nair & Weichert *Bioorg Chem* **9** 423 1980, Löfgren et al. *Acta Chem Scand* **7** 225 1953, UV: Brown & Weliky *J Biol Chem* **204** 1019 1953, *Beilstein* **26** III/IV 1740.]

**Puromycin dihydrochloride** (*O*-methyl-L-tyrosine[*N*<sup>6</sup>,*N*<sup>6</sup>-dimethylaminoadenosin-3'-yl-amide]) [58-58-2/ **M 616.5**, **m 174°**,  $[\alpha]_D^{25} -11^\circ$  (free base in EtOH), **pK<sub>1</sub> 6.8**, **pK<sub>2</sub> 7.2**. Puromycin dihydrochloride is purified by recrystallisation from H<sub>2</sub>O. The *free base*, [58-60-6] **M 294.3**, has **m 175.5-177°** (172-173°) (from H<sub>2</sub>O). The *sulfate* has **m 180-187°** dec (from H<sub>2</sub>O), and the *picrate monohydrate* has **m 146-149°** (from H<sub>2</sub>O). [Baker et al. *J Am Chem Soc* **77** 1 1955, Fryth et al. *J Am Chem Soc* **80** 3736 1958.] It is an inhibitor of aminopeptidase and terminates protein synthesis [Reboud et al. *Biochemistry* **20** 5281 1981]. [*Beilstein* **26** III/IV 3704.]

**Pyridoxal hydrochloride** [65-22-5] **M 203.6**, **m 176-180°(dec)**, **pK<sub>1</sub><sup>20</sup> 4.23** (3-OH), **pK<sub>2</sub><sup>20</sup> 8.7** (Pyridinium<sup>+</sup>), **pK<sub>3</sub><sup>20</sup> 13.04** (CH<sub>2</sub>OH?). Dissolve it in water and adjust the pH to 6 with NaOH. Set aside overnight to crystallise. The crystals are washed with cold water, dried in a vacuum desiccator over P<sub>2</sub>O<sub>5</sub>, and stored in a brown bottle at room temperature. The *free base* is then converted to the hydrochloride with one equivalent of HCl. [Fleck & Alberty *J Phys Chem* **66** 1678 1962, *Beilstein* **21/13** V 44.]

**Pyridoxal-5'-phosphate monohydrate** (PLP, codecarboxylase) [54-47-7] **M 265.2**, **pK<sub>1</sub><sup>25</sup> <2.5** (PO<sub>4</sub>), **pK<sub>2</sub><sup>25</sup> 4.14** (3-OH), **pK<sub>3</sub><sup>25</sup> 6.20** (PO<sub>4</sub>), **pK<sub>4</sub><sup>25</sup> 8.69** (pyridinium<sup>+</sup>). PLP has been purified by dissolving 2g in H<sub>2</sub>O (10-15ml, in a dialysis bag a third full) and dialysing with gentle stirring against 1L of H<sub>2</sub>O (+ two drops of toluene) for 15 hours in a cold room. The dialysate is evaporated to 80-100ml, then lyophilised. Lemon yellow microscopic needles of the monohydrate remain when all the ice crystals have been removed. The purity is checked by paper chromatography (in EtOH or *n*-PrOH/NH<sub>3</sub>) and the spot(s) visualised under UV light after reaction with a spray of *p*-phenylene diamine, NH<sub>3</sub> and molybdate. Solutions stored in a freezer are 2-3% hydrolysed in 3 weeks. At 25°, only 4-6% hydrolysis occurs even in N NaOH or HCl, and 2% is hydrolysed at 37° in 1 day—but is complete at 100° in 4 hours. It is best stored as a dry solid at -20°. In aqueous acid the solution is colourless but is yellow in alkaline solutions. It has UV with  $\lambda_{\max}$  at 305nm ( $\epsilon$  1100) and 380nm ( $\epsilon$  6550) in 0.1 N NaOH; 330nm ( $\epsilon$  2450) and 388nm ( $\epsilon$  4,900) in 0.05M phosphate buffer pH 7.0 and 295nm ( $\epsilon$  6700) in 0.1N HCl. [Peterson et al. *Biochemical Preparations* **3** 34, 119 1953.] The *oxime* decomposes at 229-230° and is practically insoluble in H<sub>2</sub>O, EtOH and Et<sub>2</sub>O. The *O-methylloxime* decomposes at 212-213°. [Heyl et al. *J Am Chem Soc* **73** 3430 1951.] It has also been purified by column chromatography through Amberlite IRC-50 (H<sup>+</sup>) [Peterson & Sober *J Am Chem Soc* **76** 169 1954]. [*Beilstein* **21/13** V 46.]

**Pyridoxamine hydrochloride** [5103-96-8, 524-36-7 (*free base*)] **M 241.2**, **m 226-227°(dec)**, **pK<sub>1</sub><sup>25</sup> 3.54** (3-OH), **pK<sub>2</sub><sup>25</sup> 8.21** (ring N<sup>+</sup>), **pK<sub>3</sub><sup>25</sup> 10.63** (NH<sub>2</sub>). The amine salt is recrystallised from hot MeOH. The *free base* crystallises from EtOH with **m 193-193.5°** [Harris et al. *J Biol Chem* **154** 315 1944, *J Am Chem Soc* **66** 2088 1944]. [*Beilstein* **22** IV 6064, **22/12** V 324.]

**Quizalofop ethyl** {[ethyl 2-[4-(6-chloro-2-quinoxalinyloxy)phenoxy]propionate} [(±) 76578-14-8; R- 100646-51-3, 100646-52-4, 100760-08-5, 100760-10-9] **M 372.8**, **m 92-93°**, **m 220°/0.2mm**. This (±)-herbicide forms white crystals from Me<sub>2</sub>CO/EtOH and sublimes *in vacuo*. Large quantities can be distilled at high vacuum. Its solubilities at 20° in g/10ml are 0.09 (EtOH), 1.1 (Me<sub>2</sub>CO), 1.2 (xylene), 2.9 (\*C<sub>6</sub>H<sub>6</sub>), and is nearly insoluble in H<sub>2</sub>O (0.3mg/l). The (±)-acid has [95977-28-9]. The *R-ester enantiomer* has **m 76-77°** (pale brown crystals from EtOH),  $[\alpha]_D^{20} +35.9^\circ$  (EtOH). It is the more active fatty acid synthase inhibitor (designated DPX-Y6202) used to control grassy weeds in broadleaf crops. The *R-acid* has [94051-08-8].

**Resorufin** (7-hydroxy-3H-phenoxazin-3-one Na salt) [34994-50-8] **M 252.2**, **decomposes on heating**, **pK<sup>25</sup> 5.8** (H<sub>2</sub>O), **6.10** (50% aqueous MeOH). The free acid is purified by dissolving it in conc H<sub>2</sub>SO<sub>4</sub> which on dilution precipitates resorufin and is filtered off, washed well with hot H<sub>2</sub>O and dried. [Eichler *J Prakt Chem* **139** 113 1934.] Detailed UV and IR studies were reported by Musso & Mattheis *Chem Ber* **90** 1814 1957]. [pKa: Musso & Rathjen *Chem Ber* **92** 751 1959, *Beilstein* **27** 1V 2263.]

The **monoacetate** [1152-14-3] **M 252.2**, **m 223°** is a fluorogenic substrate used for hydrolytic enzymes, and its fluorescence has  $E_{em}$  593nm ( $E_{ex}$  500nm, in 0.1M phosphate pH 8.0 with lipase) [Kitson & Kitson *Biochem J*

322 701 1997]. [*Beilstein* 27 IV 2263.]

The **O<sup>7</sup>-methylresorufin (7-methoxy-3*H*-phenoxazin-3-one)** [5735-89-3] **M 227.2, m >220°** has fluorescence at  $\lambda_{\text{em}}$  585nm ( $\lambda_{\text{ext}}$  571 nm in deacylase solution) and is used to differentiate isoenzymes of cytochrome P-450. It is insoluble in H<sub>2</sub>O and dilute alkali but is soluble in EtOH and CHCl<sub>3</sub> to give an orange yellow colour. [Kehrmann *Justus Liebigs Ann Chem* 372 352 1910, *Beilstein* 27 II 108.]

**Riboflavin.** See vitamin B<sub>2</sub> or lactoflavin below.

**Riboflavin-5'-phosphate (Na salt, 2H<sub>2</sub>O).** See flavin mononucleotide (FMN) above.

**Rifampicin (Rifampin)** [13292-46-1] **M 823.0, m 183-185°, pK<sub>1</sub> 1.7, pK<sub>2</sub> 7.9.** This macrolide antibiotic crystallises from Me<sub>2</sub>CO in red-orange plates. It has UV with  $\lambda_{\text{max}}$  at 237, 255, 334, and 475nm ( $\epsilon$  33,200, 32,100, 27,000 and 15,400) at pH 7.38. It is stable in Me<sub>2</sub>SO and H<sub>2</sub>O and is freely soluble in most organic solvents, but slightly soluble in H<sub>2</sub>O at pH <6. [Binda et al. *Arzneim.-Forsch* 21 1907 1971.] It inhibits cellular RNA synthesis without affecting DNA [Calvori et al. *Nature* 207 417 1965].

**Rifamycin B** [13929-35-6] **M 755.8, m 300° (darkening at 160-164°),  $[\alpha]_{\text{D}}^{20}$  -11° (MeOH), pK<sub>1</sub> 2.60, pK<sub>2</sub> 7.76.** Rifamycin B forms yellow needles from \*C<sub>6</sub>H<sub>6</sub>. Its solubilities are: H<sub>2</sub>O (0.027%), MeOH (2.62%) and EtOH (0.44%). It has UV with  $\lambda_{\text{max}}$  at 223, 304 and 245nm ( $A_{1\text{cm}}^{1\%}$  555, 275 and 220). [Oppolzer & Prelog *Helv Chim Acta* 56 2287 1973, Oppolzer et al. *Experientia* 20 336 1964, X-ray: Brufani et al. *Experientia* 20 339 1964.]

**Rifamycin SV** [6998-60-3] **M 697.8, m 300°(darkening >140°),  $[\alpha]_{\text{D}}^{20}$  -4° (MeOH), pK<sub>Est</sub> ~7.8.** Rifamycin SV gives yellow-orange crystals from Et<sub>2</sub>O/petroleum ether or aqueous EtOH, is very soluble in MeOH, EtOH, Me<sub>2</sub>CO and EtOAc, and is less soluble in Et<sub>2</sub>O and HCO<sub>3</sub><sup>-</sup>, but slightly soluble in H<sub>2</sub>O and petroleum ether. Its UV has  $\lambda_{\text{max}}$  at 223, 314 and 445nm ( $A_{1\text{cm}}^{1\%}$  586, 322 and 204) in phosphate buffer pH 7. [NMR: Bergamini & Fowst *Arzneim.-Forsch* 15 951 1965.]

**(-)-Scopolamine hydrobromide 3H<sub>2</sub>O (6 $\beta$ ,7 $\beta$ -epoxy-3 $\alpha$ -tropanyl *S*(-)-tropate HBr, hyoscine HBr)** [114-49-8] **M 438.3, m 193-194°, 195°, 195-199°,  $[\alpha]_{\text{D}}^{25}$  -25° (c 5, H<sub>2</sub>O), pK<sup>20</sup> 8.15.** The hydrobromide is recrystallised from Me<sub>2</sub>CO, H<sub>2</sub>O or EtOH/Et<sub>2</sub>O and dried. It is soluble in H<sub>2</sub>O (60%) and EtOH (5%) but insoluble in Et<sub>2</sub>O and slightly in CHCl<sub>3</sub>. The *hydrochloride* has **m 300°** (from Me<sub>2</sub>CO). The *free base* is a viscous liquid which forms a crystalline *hydrate* with **m 59°** and  $[\alpha]_{\text{D}}^{20}$  -28° (c 2.7, H<sub>2</sub>O). It hydrolyses in dilute acid or base. [Meinwald *J Chem Soc* 712 1953, Fodor *Tetrahedron* 1 86 1957, *Beilstein* 6 III 4185.]

**Serotonin hydrochloride (5-HT, 3-[2-aminoethyl]-5-hydroxyindole HCl)** [153-98-0] **M 212.7, m 167-168°, 178-180°, pK<sub>1</sub><sup>25</sup> 4.9, pK<sub>2</sub><sup>25</sup> 9.8 (10.0, NH<sub>2</sub>), pK<sub>3</sub><sup>25</sup> 11.1 (5-OH), pK<sub>4</sub><sup>25</sup> 18.25 (acidic indole NH).** 5-HT is purified by recrystallisation from EtOH/Et<sub>2</sub>O or Et<sub>2</sub>O to give the *hygroscopic* salt. Store it in the dark as it is light sensitive. The *free base* has **m 84-86°** (from Et<sub>2</sub>O). The *5-benzyloxy* derivative has **m 84-86°** (from Et<sub>2</sub>O). [Ek & Witkop *J Am Chem Soc* 76 5579 1954, Hamlin & Fischer *J Am Chem Soc* 73 5007 1951.] The *picrate* *1H<sub>2</sub>O* has **m 196-197.5°** (dec with sintering at 160-165°) after crystallisation from Et<sub>2</sub>O. Serotonin is a natural neurotransmitter [Chuang *Life Sci* 41 1051 1987]. [*Beilstein* 22/12 V 16.]

**Spectinomycin dihydrochloride pentahydrate (Actinospectacin)** [21736-83-4] **M 495.3, m 205-207°(dec),  $[\alpha]_{\text{D}}^{20}$  +14.8° (c 0.4, H<sub>2</sub>O), pK<sub>1</sub> 6.95, pK<sub>2</sub> 8.70.** The salt is purified by recrystallisation from aqueous Me<sub>2</sub>CO and is soluble in H<sub>2</sub>O, MeOH and dilute acid and base, but only slightly soluble in Me<sub>2</sub>CO, EtOH, CHCl<sub>3</sub> and \*C<sub>6</sub>H<sub>6</sub>. The *free base* is an amorphous solid, **m 184-194°** with  $[\alpha]_{\text{D}}^{20}$  -20° (H<sub>2</sub>O). [Wiley et al. *J Am Chem Soc* 93 2652 1963, X-ray: Cochran et al. *J Chem Soc Chem Commun* 494 1972.] It is an aminoglycoside antibiotic which interacts with 16S ribosomal RNA [Moazet & Noller *Nature* 327 389 1987] and is used for the treatment of gonorrhoea [Rinehart *J Infect Dis* 119 345 1969].

**D-Sphingosine (2*S*,3*S*-D-erythro-2-amino-octadec-4*t*-ene-1,3-diol from bovine brain)** [123-78-4] **M 299.5, m 79-82°, 82° 82.5° (softens at ~70°),  $[\alpha]_{\text{D}}^{22}$  -3.4° (c 2, CHCl<sub>3</sub>), pK<sub>Est</sub> ~8.8.** D-Sphingosine is purified by re-

crystallisation from EtOAc, Et<sub>2</sub>O or petroleum ether (60–80°). It is insoluble in H<sub>2</sub>O but is soluble in Me<sub>2</sub>CO, EtOH and MeOH. It has IR bands at 1590 and 875 cm<sup>-1</sup>, and is characterised as the *tribenzoate* **m** 122–123° (from 95% EtOH). [Tipton *Biochemical Preparations* **9** 127 1962.]

**Squalene** precursor of cholesterol. See “Carotenoids” in this chapter.

**Sterigmatocystin** (3a,12c-dihydro-8-hydroxy-6-methoxy-3*H*-furo[3',2':4,5]furo[2,3-*c*]xanthen-7-one) [10048-13-2] **M** 324.3, **m** 246°, 247–248°, [α]<sub>D</sub><sup>20</sup> -398° (c 0.1, CHCl<sub>3</sub>), pK<sub>Est</sub> ~ 8.0. It crystallises from amyl acetate, Me<sub>2</sub>CO or EtOH and sublimes *in vacuo*. It has UV with λ<sub>max</sub> at 208, 235, 249 and 329nm (log ε 4.28, 4.39, 4.44 and 4.12). [UV: Bullock et al. *J Chem Soc* 4179, 1962, UV, IR: Holker & Mulheim *J Chem Soc, Chem Commun* 1576, 1576 1968, Birkinshaw & Hammady *Biochem J* **65** 162 1957.] This mycotoxin induces bone marrow changes in mice [Curry et al. *Mutation Res* **137** 111 1984]. [Beilstein **19/10** V 575.]

**Stigmatellin A** (2-[4,6-dimethoxy-3,5,11-trimethyltridecatri-7*t*,9*t*,11*t*-enyl]-8-hydroxy-5,7-dimethoxy-3-methyl-4*H*-1-benzopyran-4-one) [91682-96-1] **M** 514.6, **m** 128–130°, [α]<sub>D</sub><sup>20</sup> +38.5° (c 2.3, MeOH), pK<sub>Est</sub> ~7 (phenolic OH). Stigmatellin A is stable in aqueous solution at neutral pH but decomposes at pH <5. It is purified by recrystallisation from toluene/hexane. It has UV with λ<sub>max</sub> nm (ε) at 248sh (4,100), 258 (59,500), 267 (65,500), 279 (1400) and 335 (5200) in MeOH; 249sh (45,600), 258 (60,000), 268 (72,700), 277 (54,100), 320 (2,500) and 370 (3000) in MeOH + 1 drop of N KOH; 243sh (29,300), 264 (63,200), 274 (64,100), 283sh (45,800), 329 (4800) and 420 (21,000) in MeOH + 6N HCl; and IR (CHCl<sub>3</sub>) with ν<sub>max</sub> at 3550m, 1645chs, 1635ss, 1620ss, 1590s, 1510m and 905m cm<sup>-1</sup>. It gives colour reactions at 110° with vanillin/H<sub>2</sub>SO<sub>4</sub> (grey), Ce(IV)/(NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub> (yellow) and phosphomolybdate (blue-grey). [Höfle et al. *Justus Liebigs Ann Chem* 1882 1984.] It inhibits electron transport [Jagow & Link *Methods Enzymol* **126** 253 1986, Robertson et al. *Biochemistry* **32** 1310 1933], and has antibiotic properties [Kunze et al. *J Antibiot* **37** 454 1984]. The 7*t*,9*t*,11*c*-isomer is *Stigmatellin B*.

**Streptomycin sulfate** [3810-74-0] **M** 1457.4, [α]<sub>D</sub><sup>20</sup> -84.3° (c 3, H<sub>2</sub>O), pK<sub>Est(1)</sub> ~ 9.5 (MeNH), pK<sub>Est(2,3)</sub> ~ 13.4 (guanidino). The sulfate is recrystallised from H<sub>2</sub>O/EtOH, washed with a little EtOH, Et<sub>2</sub>O and dried in a vacuum. [UV and IR: Grove & Randall *Antibiotics Monographs* **2** 163 1855, Heuser et al. *J Am Chem Soc* **75** 4013 1953, Kuehl et al. *J Am Chem Soc* **68** 1460 1946, Regna et al. *J Biol Chem* **165** 631 1946.] During protein synthesis it inhibits initiation and causes misreading of mRNA [Zierhut et al. *Eur J Biochem* **98** 577 1979, Chandra & Gray *Methods Enzymol* **184** 70 1990]. [Beilstein **18/11** V 82.]

**Streptonigrin (nigrin, 5-amino-6-[7-amino-5,8-dihydro-6-methoxy-5,8-dioxo-2-quinolinyl]-4-[2-hydroxy-3,4-dimethoxyphenyl]-3-methyl-2-pyridinecarboxylic acid)** [3930-19-6] **M** 506.5, **m** 262–263°, 275°(dec), pK<sub>25</sub> 6.3 (1:1 aqueous dioxane). Streptonigrin is purified by TLC on pH 7-buffered silica gel plates (made from a slurry of Silica Gel 60 and 400ml of 0.05M phosphate buffer pH 7.0) and eluted with 5% MeOH/CHCl<sub>3</sub>. Material from the extracted band recrystallises from Me<sub>2</sub>CO or dioxane as almost black plates or needles. It is soluble in pyridine, Me<sub>2</sub>NCHO, aqueous NaHCO<sub>3</sub> (some dec), and slightly soluble in MeOH, EtOH, EtOAc and H<sub>2</sub>O. It has UV with λ<sub>max</sub> at 248, 375–380nm (ε 38,400 and 17,400). [Weinreb et al. *J Am Chem Soc* **104** 536 1982, Rao et al. *J Am Chem Soc* **85** 2532 1963.] It is antineoplastic and causes severe bone marrow depression [Wilson et al. *Antibiot Chemother* **11** 147 1961].

**Succinyl coenzyme A trisodium salt** [108347-97-3] **M** 933.5. If it should be purified further, then it should be dissolved in H<sub>2</sub>O (0.05g/ml) adjusted to pH 1 with 2M H<sub>2</sub>SO<sub>4</sub> and extracted several times with Et<sub>2</sub>O. Excess Et<sub>2</sub>O is removed from the aqueous layer by bubbling N<sub>2</sub> through it and is stored frozen at pH 1. When required, the pH should be adjusted to 7 with dilute NaOH and used within 2 weeks (samples should be frozen). Succinyl coenzyme A is estimated by the hydroxamic acid method [Hersch & Jencks *J Biol Chem* **242** 3468 1967]. It is more stable in acidic than in neutral aqueous solutions, but neutral solutions can be stored at -15° with negligible decomposition. [Jordan & Laghai-Newton *Methods Enzymol* **123** 435 1986, Beilstein **26** III/IV 3666.]

**Terramycin (oxytetracycline)** [79-57-2; 6153-64-6 (2H<sub>2</sub>O)] **M** 460.4 (anhydrous), 496.5 (2H<sub>2</sub>O), sinters at 182°, melts at 184–185°(dec), [α]<sub>D</sub><sup>20</sup> -196.6° (equilibrium in 0.1M HCl), -2.1°(equilibrium in 0.1M NaOH).

Terramycin crystallises (as dihydrate) from water or aqueous EtOH and is soluble in MeOH, EtOH, Me<sub>2</sub>CO and H<sub>2</sub>O (0.25mg/ml at 25°) but insoluble in Et<sub>2</sub>O and petroleum ether. It is amphoteric, and an aqueous solution has pH 2.0-5.0. It has UV with  $\lambda_{\max}$  at 247, 275 and 353nm in 0.1 M phosphate (pH 4.5). [Finlay et al. *Science* **111** 85 1951.] The *hydrochloride*, [2058-46-0] M 496.9 [Beilstein **14** IV 2630], crystallises from MeOH in needles and from H<sub>2</sub>O at 50° it forms plates. Terramycin has also been purified *via* the *hydrochloride* by dissolving it in H<sub>2</sub>O, adjusting to pH 6, and the solid is filtered off after 1 hour. The crystals of the *dihydrate* are dried to constant weight in vacuum/CaCl<sub>2</sub>/25°. Drying at 60° *in vacuo* gives the *anhydrous base* m 184.5-185.5° (sintering at 180°). The *dihydrate* has m 181-182°. Its optical rotation in MeOH decreases from  $[\alpha]_{\text{D}}^{25} +2.6^{\circ}$  (c 0.5%) to  $[\alpha]_{\text{D}}^{25} +11.3^{\circ}$  after standing for 16 hours. It forms a *sodium salt* and a *CaCl<sub>2</sub> complex*. [Regna et al. *J Am Chem Soc* **73** 4211 1951, Beilstein **14** IV 2633.]

**Tetracycline** [60-54-8] M 444.4, m 172-174°(dec),  $[\alpha]_{\text{D}}^{20} -291^{\circ}$  (c 1, MeOH), pK<sub>1</sub><sup>25</sup> 3.30, pK<sub>2</sub><sup>25</sup> 7.68, pK<sub>3</sub><sup>25</sup> 9.69. Tetracycline crystallises from toluene or aqueous MeOH as the *trihydrate*. [Stephen et al. *J Am Chem Soc* **76** 3568 1954, Beilstein **14** IV 2625.]

**Tetracycline hydrochloride** [64-75-5] M 480.9, m 214°(dec), 215-220°,  $[\alpha]_{\text{D}}^{25} -258^{\circ}$  (c 0.5, 0.1N HCl),  $[\alpha]_{\text{D}}^{20} -245^{\circ}$  (c 1, MeOH),  $[\alpha]_{\text{D}}^{25} -257.9^{\circ}$  (c 0.5, aqueous HCl), pK<sub>1</sub> 1.4 (enolic OH), pK<sub>2</sub> 3.30, pK<sub>3</sub> 7.68 (phenolic OH), pK<sub>4</sub> 9.27 (Me<sub>2</sub>N). The hydrochloride is recrystallised from MeOH/*n*-BuOH or *n*-BuOH/HCl. It is insoluble in Et<sub>2</sub>O and petroleum ether. It has UV with  $\lambda_{\max}$  at 270 and 366nm in MeOH. [Gottstein et al. *J Am Chem Soc* **81** 1198 1959, Conover et al. *J Am Chem Soc* **84** 3222 1962, Stephen et al. *J Am Chem Soc* **78** 4155 1956, Beilstein **14** IV 2627.]

**6R-Tetrahydro-erythro-biopterin dihydrochloride** (BH<sub>4</sub>.2HCl, 6R-2-amino-4-hydroxy-6-{{1R,2S}-1,2-dihydroxypropyl}-5,6,7,8-tetrahydropteridine 2HCl) [69056-38-8] M 316.2, m 245-246°(dec),  $[\alpha]_{\text{D}}^{25} -6.8^{\circ}$  (c 0.67, 0.1N HCl), pK<sub>1</sub> 1.37 (pyrimidine<sup>+</sup>), pK<sub>2</sub> 5.6 (5-NH<sup>+</sup>), pK<sub>3</sub> 10.6 (acidic, 3NH). Recrystallisation of BH<sub>4</sub>.2HCl from HCl enriches BH<sub>4</sub> in the natural 6R isomer. Dissolve the salt (~6g) in conc HCl (15ml) under gentle warming, then add EtOH (30ml) dropwise, chill and collect the colourless needles (67%, up to 99% if the mother liquors are concentrated), and dry it *in vacuo* immediately over P<sub>2</sub>O<sub>5</sub> and KOH. It is stable indefinitely at -20° in a dry atmosphere. Better store it in sealed ampoules under dry N<sub>2</sub>. It can be recrystallised from 6N aqueous HCl. It has UV with  $\lambda_{\max}$  (in 2N HCl) at 264nm ( $\epsilon$  16,770; pH 3.5 phosphate buffer) 265nm ( $\epsilon$  13,900); (in pH 7.6) at 297nm ( $\epsilon$  9500) and 260nm sh ( $\epsilon$  4690). It has been separated from the 6S-isomer by HPLC on a Partisil-10SCX column using 30mM ammonium phosphate buffer (pH 3.0) containing 3mM NaHSO<sub>3</sub> (2ml/minute flow rate; 275nm detector) with retention times of 5.87 minutes (6R) and 8.45 minutes (6S). It is stable in acidic solutions and can be stored for extended periods at -20° in 0.04M HCl. Above pH 7 the neutral species are obtained, and these are readily oxidised by the oxygen in the solvent to the *quinonoid species*, and then further oxidation and degradation occurs at room temperature. These changes are slower at 0°. The *sulfate salt* can be obtained by recrystallisation from 2M H<sub>2</sub>SO<sub>4</sub> and is less soluble in water than the hydrochloride salt. The 6R-2,5,1',2'-tetraacetylbiopterin derivative has m 292°(dec) after recrystallisation from MeOH (100 parts) and  $[\alpha]_{\text{D}}^{20} -144^{\circ}$  (c 0.5, CHCl<sub>3</sub>),  $[\alpha]_{\text{D}}^{20} +12.8^{\circ}$  (c 0.39, Me<sub>2</sub>SO). [NMR, UV: Matsuura et al. *Heterocycles* **23** 3115 1985, Viscontini et al. *Helv Chim Acta* **62** 2577 1979, Armarego et al. *Aust J Chem* **37** 355 1984, Beilstein **26** III/IV 4032.]

**Tetrahydrofolic acid dihydrochloride 2H<sub>2</sub>O** (THFA, 6S- or 6RS- 5,6,7,8-tetrahydrofolic acid 2HCl 2H<sub>2</sub>O, 5,6,7,8-tetrahydropteroyl-L-glutamic acid 2HCl 2H<sub>2</sub>O) [135-16-0] M 544.4, m >200°(dec),  $[\alpha]_{\text{D}}^{27} +16.9^{\circ}$  (H<sub>2</sub>O pH 7.0 + 2-mercaptoethanol), pK<sub>1</sub> 1.7 (pyrimidine N<sup>+</sup>), pK<sub>2</sub> 2.4 (10N<sup>+</sup>), pK<sub>3</sub> 3.5 ( $\alpha$ -CO<sub>2</sub>H), pK<sub>4</sub> 4.9 ( $\gamma$ -CO<sub>2</sub>H), pK<sub>5</sub> 5.6 (5-NH<sup>+</sup>), pK<sub>6</sub> 10.4 (acidic, 3NH). Very high quality material is now available commercially, and it should be a white powder. It can be dried over P<sub>2</sub>O<sub>5</sub> in a vacuum desiccator and stored in weighed aliquots in sealed ampoules. It is stable at room temperature in sealed ampoules for many months and for much more extended periods at -10°. When moist, it is extremely sensitive to air whereby it oxidises to the yellow 7,8-dihydro derivative. In solution it turns yellow in colour as it oxidises, and then particularly in the presence of acids it turns dark reddish brown in colour. Hence aqueous solutions should be frozen immediately when not in use. It is always advisable to add 2-mercaptoethanol (if it does not interfere with the procedure for which it is used) which stabilises it by depleting the solution of O<sub>2</sub>. The *sulfate salt* is more stable but is much less soluble. The best way to prepare standard solutions of this acid is to dissolve it in the desired buffer and

estimate the concentration by UV absorption in pH 7 buffer at 297nm ( $\epsilon$  22,000 M<sup>-1</sup>cm<sup>-1</sup>). If a sample is suspect, it is not advisable to purify it because it is likely to deteriorate further as “dry box” conditions are necessary. Either a new sample is purchased or one is freshly prepared from folic acid. It has the above pKa values. [Hafeti et al. *Biochemical Preparations* 7 89 1960, UV: Mathews & Huennekens *J Biol Chem* 235 3304 1960, Osborn & Huennekens *J Biol Chem* 233 969 1958, O’Dell et al. *J Am Chem Soc* 69 250 1947, Blakley *Biochem J* 65 331 1957, Asahi *Yakugaku Zasshi (J Pharm Soc Jpn)* 79 1548 1959, Beilstein 26 III/IV 3879.]

**5,6,7,8-Tetrahydropterin sulfate (2-amino-5,6,7,8-tetrahydropteridin-4-one H<sub>2</sub>SO<sub>4</sub>)** [20350-44-1] **M 265, m >200°(dec), pK<sub>1</sub><sup>25</sup> 1.3 (pyrimidine<sup>+</sup>), pK<sub>2</sub><sup>25</sup> 5.6 (5-NH<sup>+</sup>), pK<sub>3</sub><sup>25</sup> 10.6 (acidic, 3NH).** If its colour has become strongly violet, then it will need to be reduced again. It is best to check the UV absorption in N HCl where it has a peak at ~265nm which drops sharply to zero having no absorption at ~340nm. The presence of absorption at 340nm indicates oxidation to quinonoid or 7,8-dihydropterin. If the absorption is weak, then dissolve it in the minimum volume of anhydrous trifluoroacetic acid (fume hood), add charcoal, filter, then add two drops of N H<sub>2</sub>SO<sub>4</sub> followed by dry Et<sub>2</sub>O at 0°, allow the white tetrahydro salt to settle, collect, and wash it with dry Et<sub>2</sub>O, by centrifugation. Dry the residue *in vacuo* over P<sub>2</sub>O<sub>5</sub> and KOH. Store it in aliquots in the dark at <0°. It has UV  $\lambda_{\max}$  at 265nm ( $\epsilon$  16,980) at pH -1.0 (dication); 219nm ( $\epsilon$  23,440) and 266nm ( $\epsilon$  12,880) at pH 3.5 (monocation); 220nm ( $\epsilon$  18,620), [260nm ( $\epsilon$  4270)sh] and 299nm ( $\epsilon$  9330) at pH 8.0 (neutral species); and 218nm ( $\epsilon$  10,000), [240nm ( $\epsilon$  5500)sh] and 287nm ( $\epsilon$  5500) at pH 13 (anion). [Blakley *Biochem J* 72 707 1959, Asahi *Yakugaku Zasshi (J Pharm Soc Jpn)* 79 1557 1959, Pfeleiderer in *Pterins and Folate* (Benkovic and Blakley Eds) J Wiley Vol 2 p 97 1985.]

**Thiamine monophosphate chloride 2H<sub>2</sub>O (Aneurine monophosphate chloride)** [532-40-1] **M 416.8, m 193°(dec), 200°(dec), 200-203°(dec), pK<sub>1</sub> 2.40, pK<sub>2</sub> 4.80, pK<sub>3</sub> 6.27, pK<sub>4</sub> 9.65, pK<sub>5</sub> 10.20.** Purify it by recrystallisation from aqueous HCl, EtOH slightly acidified with HCl, EtOH/Me<sub>2</sub>CO, H<sub>2</sub>O, or H<sub>2</sub>O/EtOH/Et<sub>2</sub>O. Dissolve it in a small volume of H<sub>2</sub>O and mix it with EtOH/Me<sub>2</sub>CO (1:1) to give the HCl.H<sub>2</sub>O as crystals. Filter it off, wash it with Et<sub>2</sub>O and dry it in a vacuum. The *chloride hydrochloride*, **m 215-217°(dec)** is obtained when it is crystallised from aqueous HCl. [Wenz et al. *Justus Liebigs Ann Chem* 618 2280 1958, Viscontini et al. *Helv Chim Acta* 34 1388 1951, Leichssenring & Schmidt *Chem Ber* 95 767 1962, McCormick & Wright *Methods Enzymol* 18A 141, 147 1970, Beilstein 27 III/IV 1766.]

**Thiamphenicol (Thycymetin, 1R,2R-2-[2,2-dichloroacetyl-amino]-1-[4-methanesulfonyl-phenyl]-propan-1,3-diol)** [15318-45-3 (*D-threo*), 90-91-5] **M 356.2, m 163-166°, 165.2-165.6°, 165-166°,  $[\alpha]_{\text{D}}^{25}$  +15.6° (c 2, EtOH), pK<sup>25</sup> 7.2.** Recrystallise thiamphenicol from H<sub>2</sub>O or CHCl<sub>3</sub>. The UV has  $\lambda_{\max}$  at 224, 266 and 274nm ( $\epsilon$  13,700, 800 and 700) in 95% EtOH. The *1S,2S-isomer* [14786-51-7] has **m 164.3-166.3°** (from H<sub>2</sub>O/EtOAc/petroleum ether) and  $[\alpha]_{\text{D}}^{25}$  -12.6° (c 1, EtOH); and the *racemate 1RS,2RS-Racefenical* [847-25-6] has **m 181-183° (dec)** from CHCl<sub>3</sub>/EtOAc/petroleum ether. [Cutler et al. *J Am Chem Soc* 74 5475, 5482 1952, UV: Nachod & Cutler *J Am Chem Soc* 74 1291 1952, Suter et al. *J Am Chem Soc* 75 4330 1953, Cutler et al. *J Am Pharm Assoc* 43 687 1954, Beilstein 13 IV 2957.]

**$\epsilon$ -[2-(4-Thiazolidinone)]hexanoic acid (Mycobacidin, Acidomycin, 6[4-oxothiazolidin-2-yl]hexanoic acid)** [539-35-5] **M 215.3, m 140°, pK<sup>25</sup> 5.1.** The *dl-form* is *dimorphic*; it crystallises from CHCl<sub>3</sub> with **m 116-117°**, and from H<sub>2</sub>O with **m 123°**. The *l(-)-enantiomer* (from Actinomyces) crystallises from H<sub>2</sub>O, MeOH (**m 139-140°**), aqueous EtOH (**m 140-141°**) or EtOAc, and has  $[\alpha]_{\text{D}}^{20}$  -54° (c 1, MeOH). The *d(+)-enantiomer* (from optical resolution of the brucine salt) has **m 138-139°** (from H<sub>2</sub>O) and  $[\alpha]_{\text{D}}^{25}$  +57° (c 1, MeOH). The optically active acids racemise in hot alkali. [McLamore et al. *J Am Chem Soc* 75 105 1953, Beilstein 27 III/IV 4281.]

**6-Thioguanosine (2-amino-6-mercapto-9- $\beta$ -D-ribofuranosylpurine)** [85-31-4] **M 299.3, m 224-227°(dec), 230-231°(dec),  $[\alpha]_{\text{D}}^{20}$  -64° (c 1.3, 0.1N NaOH), pK<sup>25</sup> 8.33.** Thioguanosine is crystallised (as *hemihydrate*) from hot H<sub>2</sub>O (charcoal) and cooled slowly to give tapered prisms. It also crystallises by dissolving in dilute NH<sub>3</sub> and acidifying with acetic acid, and then recrystallising from H<sub>2</sub>O. Its UV (pH 4-6) has  $\lambda_{\max}$  at 257nm ( $\epsilon$  8,820) and 342nm ( $\epsilon$  24,800), and at pH 10.4-12.0 it has  $\lambda_{\max}$  at 252nm ( $\epsilon$  14,700) and 319.5nm ( $\epsilon$  21,000). [Fox et al. *J Am Chem Soc* 80 1667 1958, Beilstein 26 III/IV 3927.]

***dl*- $\alpha$ -Tocopherol (see vitamin E)** [10191-41-0] **M 430.7,  $A_{1\text{cm}}^{1\%}$  74.2 at 292 nm in MeOH.** Dissolve *dl*- $\alpha$ -

tocopherol in anhydrous MeOH (15ml/g) cool to  $-6^{\circ}$  for 1 hour, then chill in a Dry-ice/acetone bath; crystallisation is induced by scratching with a glass rod. The *dl*- $\alpha$ -acetate [52225-20-4] (see *DL-vitamin E acetate* below) is a viscous yellow liquid with **m**  $-7^{\circ}$ , **b**  $184^{\circ}/0.01\text{mm}$ ,  $224^{\circ}/0.3\text{mm}$ ,  $d_4^{20}$  0.953,  $n_D^{20}$  1.496. It is used as a standard for Vitamin E activity where the unit of activity is attained with 1mg of pure *dl*- $\alpha$ -acetate. [Friedrich "Vitamins" Water de Guyter Publ, Berlin 1988, *Beilstein* 17/4 V 168.]

**$\gamma$ -Tocopherol (3,4-dihydro-2,7,8-trimethyl-2-(4,8,12-trimethyltridecyl)-2H-benzopyran-6-ol)** [54-28-4] **M 416.7**, **m**  $-30^{\circ}$ , **b**  $200-210^{\circ}/0.1\text{mm}$ ,  $d_4^{20}$  0.951,  $n_D^{20}$  1.505,  $[\alpha]_D^{20}$   $-2.4^{\circ}$  (EtOH).  $\gamma$ -Tocopherol is purified by distillation at high vacuum and stored in dark ampoules under  $N_2$ . Its UV has  $\lambda_{\text{max}}$  at 298nm ( $A_{1\text{cm}}^{1\%}$  92.8). It is insoluble in  $H_2O$  but soluble in organic solvents. The *allophanate* (used for separating it from its isomers) has **m**  $136-138^{\circ}$ ,  $[\alpha]_D^{18}$   $+3.4^{\circ}$  ( $CHCl_3$ ). [Baxter et al. *J Am Chem Soc* 65 918 1943, Emerson et al. *Science* 83 421 1936, *J Biol Chem* 113 319 1936, *Beilstein* 17/4 V 158.]

**Tomatidine (5 $\alpha$ ,20 $\beta$ ,22 $\alpha$ ,25 $\beta$ ,27-azaspirostan-3 $\beta$ -ol)** [77-59-8] **M 415.7**, **m**  $202-206^{\circ}$ ,  $[\alpha]_D^{20}$   $+5.9^{\circ}$  (c 1, MeOH),  $[\alpha]_D^{20}$   $+8^{\circ}$  ( $CHCl_3$ ). Tomatidine forms plates from EtOAc. It is also purified by dissolving 80mg in  $*C_6H_6$  and applying to an  $Al_2O_3$  column (3.0g) and eluting with  $*C_6H_6$ , evaporating and recrystallising the residue three times from EtOAc. The *hydrochloride* has **m**  $265-270^{\circ}$  from EtOH and  $[\alpha]_D^{25}$   $-5^{\circ}$  (MeOH). [IR: Uhle *J Am Chem Soc* 83 1460 1961, Kessar et al. *Tetrahedron* 27 2869 1971, Schreiber & Adams *Experientia* 17 13 1961, *Beilstein* 27 III/IV 1950.]

**Tomatine (22S,25S-3 $\beta$ , $\beta$ -lycotetraosyloxy-5 $\alpha$ -spirosolan)** [17406-45-0] **M 1034.2**, **m**  $263-268^{\circ}$ (dec),  $290-291^{\circ}$ (evacuated capillary),  $283.5-287^{\circ}$ (dec),  $272-277^{\circ}$ (dec),  $300-305^{\circ}$ (dec),  $[\alpha]_D^{20}$   $-18^{\circ}$  to  $-34^{\circ}$  (c 0.55, pyridine). Tomatine is recrystallised from MeOH, EtOH, aqueous EtOH or dioxane/ $NH_3$ . It is almost insoluble in petroleum ether,  $Et_2O$  or  $H_2O$ . [Reichstein *Angew Chem* 74 887 1962, *Beilstein* 27 III/IV 1954.]

**Tubercidin (7-deazaadenosine)** [69-33-0] **M 266.3**, **m**  $247-248^{\circ}$ ,  $[\alpha]_D^{17}$   $-67^{\circ}$  (50% aqueous AcOH), **pK<sup>10</sup> 5.2-5.3**. 7-Deazaadenosine forms needles from hot  $H_2O$ . It is soluble in  $H_2O$  (0.33%), MeOH (0.5%) and EtOH (0.05%). It has UV with  $\lambda_{\text{max}}$  at 270nm ( $\epsilon$  12,100) in 0.001N NaOH. The *picrate* has **m**  $229-231^{\circ}$ (dec). [Tolman et al. *J Am Chem Soc* 91 2102 1969, Mizuno et al. *J Org Chem* 28 3329 1963, IR: Anzai et al. *J Antibiot (Japan)* [9] 10 201 1957, *Beilstein* 26 IV 1117.]

**Tunicamycin** [11089-65-9] **M ~780**, **m**  $234-235^{\circ}$ (dec),  $[\alpha]_D^{20}$   $+52^{\circ}$  (c 0.5, pyridine), **pK<sub>Est</sub> ~ 9.4**. The components of this nucleotide antibiotic from *Streptomyces* sp. are purified by recrystallising 3 times from hot glass-distilled MeOH, and the white crystals are dissolved in 25% aqueous MeOH and separated on a Partisil ODS-10 $\mu$  column (9.4 x 25 cm) [Magnum-9 Whatman] using a 260nm detector. The column is eluted with MeOH/ $H_2O$  mixture adjusted to 1:4 (v/v) then to 2:4 (v/v). The individual components are recovered and lyophilised. Ten components have been isolated, and all were active (to varying extents) depending on the lengths of the aliphatic side-chains. The mixture has UV with  $\lambda_{\text{max}}$  at 205 and 260nm ( $A_{1\text{cm}}^{1\%}$  230 and 110). It is stable in  $H_2O$  at neutral pH but unstable in acidic solution. It inhibits protein glycosylation. [Mahoney & Duskin *J Biol Chem* 254 6572 1979, Elnein *Trends Biochem Sci* 6 219 1981, Takatsuki *J Antibiot* 24 215 1971.]

**Uracil, uridine and uridine nucleotides.** These are resolved by ion-exchange chromatography with AG1 (Cl form). [Lindsay et al. *Anal Biochem* 24 506 1968.]

**Uridine 5'-(1-thio) monophosphate** [15548-52-4, 18875-72-4 (*Absolute Stereochemistry specified*)] and **Uridine 5'-( $\alpha$ -thio) diphosphate** [*RS*( $\alpha$ -P) 27988-67-6; *R*( $\alpha$ -P) 72120-52-6], **pK<sub>Est(1)</sub> ~ 6.4**, **pK<sub>Est(2)</sub> ~ 9.5**. The  $Et_3N$  salts are purified by dissolving ~4g in 500ml of  $H_2O$  (adding a drop or two of  $Et_3N$  if they do not dissolve) and chromatograph by applying to a column (3 x 30cm) of DEAE-Sephadex A-25 and eluting with 1.4L of a linear gradient of  $Et_3NH.HCO_3$  from 0.05 to 0.55M, pH 7.8 and  $4^{\circ}$ . The product elutes between 0.2-0.3M  $Et_3NH.HCO_3$ . The pooled fractions are evaporated, and the residue is twice taken up in EtOH and evaporated to dryness to remove the last traces of  $Et_3NH.HCO_3$ .  $^{31}P$  NMR:  $P_{\alpha}$  is a doublet at  $-40.81$  and  $-40.33$ , and  $P_{\beta}$  at  $7.02\text{ppm}$ ,  $J_{\alpha,\beta}$  32.96Hz. [Sheu et al. *Biochemistry* 18 5548 1979.]

**Uridylic acid (di-Na salt)** [27821-45-0] **M 368.2, m 198.5°, pK<sub>3</sub><sup>25</sup> 6.63, pK<sub>4</sub><sup>20</sup> 9.71.** Crystallise it from MeOH. It may contain some Ba salt(s); hence stir it with Amberlite IR-120 cation exchanger (25ml, wet resin, 15-50 mesh in H<sup>+</sup> form) in H<sub>2</sub>O (50ml) until the nucleotide dissolves. Filter, wash resin with H<sub>2</sub>O until the eluate is neutral. Combine the filtrate and washings, and adjust the pH to 8.0 with 2.0 M aqueous NaOH. Concentrate it *in vacuo* to a small volume (~20ml), and add Me<sub>2</sub>CO dropwise until crystallisation begins. Cool to 0°, and the shiny plates of di Na uridine-5'-phosphate dihydrate are filtered off and dried over P<sub>2</sub>O<sub>5</sub> at 25°/0.1mm for 24 hours (>76% recovery). Its UV has λ<sub>max</sub> at 262nm (ε 10,000 M<sup>-1</sup>cm<sup>-1</sup>) in 0.1 M HCl. [Boon et al. *J Chem Soc* 408 1950, Smith *Biochemical Preparations* 8 130 1960.] [*Beilstein* 24 IV 1214.]

**(+)-Usnic acid (2,6-diacetyl-7,9-dihydroxy-8,9b-dimethyldibenzofuran-1,3(2H,9bH)-dione)** [7562-61-0, 125-46-2] **M 344.3, m 201-204°, 203-206°, [α]<sub>D</sub><sup>20</sup> +630° (c 0.7, CHCl<sub>3</sub>), pK<sub>1</sub> 4.4, pK<sub>2</sub> 8.8, pK<sub>3</sub> 10.7.** This very weak acid is the natural form which is recrystallised from Me<sub>2</sub>CO, MeOH or \*C<sub>6</sub>H<sub>6</sub>. At 25° it is soluble in H<sub>2</sub>O (<0.01%), Me<sub>2</sub>CO (0.77%), EtOAc (0.88%), MeOCH<sub>2</sub>CH<sub>2</sub>OH (0.22%), and furfural (7.32%). [Curd & Robertson *J Chem Soc* 894 1937, Barton & Brunn *J Chem Soc* 603 1953, resolution: Dean et al. *J Chem Soc* 1250 1953, synthesis: Barton et al. *J Chem Soc* 538 1956, *Beilstein* 18/5 V 586.]

**(-)-Usnic acid (2,6-diacetyl-7,9-dihydroxy-8,9b-dimethyldibenzofuran-1,3(2H,9bH)-dione)** [6159-66-6, 7562-61-0] **M 344.3, m 201-204°, 204°, [α]<sub>D</sub><sup>20</sup> -495° (c 0.9, CHCl<sub>3</sub>).** Its properties are similar to those of the acid in the preceding entry. [*Beilstein* 18 II 241, 18 III/IV 3522.]

**Valinomycin (Potassium ionophore I)** [2001-95-8] **M 111.3, m 186-187°, 190°, [α]<sub>D</sub><sup>20</sup> +31.0° (c 1.6, \*C<sub>6</sub>H<sub>6</sub>).** Recrystallise valinomycin from dibutyl ether or Et<sub>2</sub>O. It is dimorphic: modification A crystallises from *n*-octane, and modification B crystallises from EtOH/H<sub>2</sub>O. It is soluble in petroleum ether, CHCl<sub>3</sub>, AcOH, BuOAc and Me<sub>2</sub>CO. [Smith et al. *J Am Chem Soc* 97 7242 1975, UV, IR and NMR see Brockmann & Schmidt-Kastner *Chem Ber* 88 57 1955, *Beilstein* 27 I 9728, 17 IV 9728.]

**(±)-Verapamil hydrochloride (5-[N-{3,4-dimethoxyphenylethyl}methylamino]-2-[3,4-dimethoxyphenyl]-2-isopropylvaleronitrile HCl)** [23313-68-0] **M 491.1, m 138.5-140.5°, 142°(dec), pK<sub>Est</sub> ~10.6.** The salt is purified by dissolving it in EtOH, filtering (if insoluble particles are present) and adding Et<sub>2</sub>O, filtering the salt, washing it with Et<sub>2</sub>O and drying it *in vacuo*. It has the following solubilities: hexane (0.001%), CH<sub>2</sub>Cl<sub>2</sub> (~10%), MeOH (~10%), EtOH (20%) and H<sub>2</sub>O (8.3%). It has UV with λ<sub>max</sub> at 232 and 278nm. The *free base* is a viscous yellow oil **b** 243-246°/0.01mm (n<sub>D</sub><sup>25</sup> 1.5448) and is almost insoluble in H<sub>2</sub>O but soluble in organic solvents. It is a Ca channel antagonist and is a coronary vasodilator. [Ramuz *Helv Chim Acta* 58 2050 1975, Harvey et al. *Biochem J* 257 95 1989.]

**Veratridine (3-veratroylveracevine)** [71-62-5] **M 673.8, m ~180° (after drying at 130°, pK 9.54 (quinolizidine N), [α]<sub>D</sub><sup>20</sup> +8.0° (c ~5, EtOH).** It is an alkaloid neurotoxin which prevents inactivation of Na<sup>+</sup> channels. Its solubility in EtOH is ~5%; and it separates as a pale yellow powder from an ethanolic solution on addition of Et<sub>2</sub>O. It forms nitrate, sulfate and perchlorate salts. [McKinney et al. *Anal Biochem* 153 33 1986, *Beilstein* 21 V/13 709.]

**Vinblastine sulfate (VLB, vincalucoblastine sulfate)** [143-67-9] **M 909.5, m 284-285°(dec), 267°(dec), [α]<sub>D</sub><sup>20</sup> -28° (MeOH), pK<sub>1</sub> 5.5, pK<sub>2</sub> 7.4.** Crystallise the sulfate from MeOH or EtOH and dry it *in vacuo* over conc H<sub>2</sub>SO<sub>4</sub>. The *free base* crystallises from EtOH or MeOH **m** 211-216° (+ 2MeOH.1 H<sub>2</sub>O) and forms a stable etherate from Et<sub>2</sub>O with **m** 201-211°, and [α]<sub>D</sub><sup>25</sup> +42° (CHCl<sub>3</sub>), and its UV has λ<sub>max</sub> at 214 and 259nm (log ε 4.73 and 4.21). The *dihydrochloride* has **m** 244-246°(dec)(MeOH). It is a monoamine oxidase B inhibitor and induces microtubule aggregation. It is an antineoplastic drug for Hodgkin's lymphoma. [Neuss et al. *J Am Chem Soc* 81 4754 1959, Jong-KeunSon et al. *J Med Chem* 33 1845 1990, Warfield & Bouck *Science* 186 1219 1974, *Beilstein* 26 III/IV 3167.]

**Vincristine sulfate (22-oxovincalucoblastine sulfate)** [2068-78-2] **M 925.1, m 218-220°, [α]<sub>D</sub><sup>25</sup> +26.2° (CH<sub>2</sub>Cl<sub>2</sub>), pK<sub>1</sub> 5.0, pK<sub>2</sub> 7.4 (in 33% Me<sub>2</sub>NCHO/H<sub>2</sub>O).** The salt is recrystallised from MeOH. It has UV with λ<sub>max</sub> at 220, 255 and 296nm (log ε 4.65, 4.21 and 4.18). It is a monoamine oxidase inhibitor and is used in

cancer research [Son et al. *J Med Chem* **33** 1845 1990, Horio et al. *Proc Natl Acad Sci USA* **85** 3580 1988].

**Viomycin sulfate (Viocin, Tuberactinomycin B)** [37883-00-4] **M 685.7, m 266°(dec)**,  $[\alpha]_D^{17}$  **-29.5° (c 1, H<sub>2</sub>O)**, **pK<sub>1</sub> 7.2 (8.2), pK<sub>2</sub> 10.3**. Viocin crystallises from H<sub>2</sub>O/EtOH and is dried in a vacuum. The dry material is *hygroscopic* and should be stored dry. The UV has  $\lambda_{\max}$  at 268 and 285nm (log  $\epsilon$  4.4 and 4.2) in H<sub>2</sub>O. [Kitigawa et al. *Chem Pharm Bull Jpn* **20** 2176 1972.] The *hydrochloride* forms *hygroscopic* plates with **m 270°(dec)**,  $[\alpha]_D^{18}$  **-16.6° (c 1, H<sub>2</sub>O)**, with  $\lambda_{\max}$  at 268nm (log  $\epsilon$  4.5) in H<sub>2</sub>O; 268nm (log  $\epsilon$  4.4) in 0.1N HCl and at 285nm (log  $\epsilon$  4.3) in 0.1N NaOH. [*Beilstein* **26** III/IV 4245.]

**Vitamin A acetate (retinyl acetate)** [127-47-9] **M 328.5, m 57°**. The acetate is separated from retinol by column chromatography, then crystallised from MeOH. [Kofler and Rubin *Vitamins and Hormones (NY)* **18** 315 1960 for purification methods]. Store it in the dark, under N<sub>2</sub> or Ar, at 0°. [*Beilstein* **6** IV 4135.]

**Vitamin A acid [Retinoic acid, 3,7-dimethyl-9-(2,6,6-trimethyl-1-cyclohexenyl)-2,4,6,8-nonatetraen-1-oic acid]** [302-79-4] **M 300.4, m 180-181°, 180-182°, pK<sub>Est</sub> ~4.2**. Purify the acid by chromatography on silicic acid columns, and eluting it with a small amount of EtOH in hexane. Also dissolve it in Et<sub>2</sub>O, wash it with H<sub>2</sub>O, dry (Na<sub>2</sub>SO<sub>4</sub>), evaporate and the solid residue is recrystallised from MeOH (0.53g /3.5ml MeOH to give 0.14g) or EtOH. It also recrystallises from *i*-PrOH, or as the *methyl ester* from MeOH. Its UV(MeOH) has  $\lambda_{\max}$  at 351nm ( $\epsilon$  45,000). *9-Cis*-acid forms yellow needles from EtOH, with **m 189-190°**, and its UV(MeOH) has  $\lambda_{\max}$  at 343nm ( $\epsilon$  36,500); the *13-cis*-acid forms red-orange plates from *i*-PrOH with **m 174-175°**, and its UV has  $\lambda_{\max}$  at 345nm ( $\epsilon$  39,800). Store it in the dark, in an inert atmosphere, at 0° [Robeson et al. *J Am Chem Soc* **77** 4111 1955]. [*Beilstein* **9** IV 2387.]

**Vitamin A alcohol (retinol)** [68-26-8] **M 286.5, A<sub>1cm</sub><sup>1%</sup> (max) (all-trans) 1, 832 (325 nm), (13-cis) 1686 (328nm), (11-cis) 1230 319 nm), (9-cis) 1480 (323 nm), (9,13-di-cis) 1379 (324 nm), (11,13-di-cis) 908 (311 nm) in EtOH**. Purify retinol by chromatography on columns of water-deactivated alumina and elute with 3-5% acetone in hexane. Separate the isomers by TLC plates on silica gel G, developed with petroleum ether (low boiling)/methyl heptanone (11:2). Store it in the dark, under N<sub>2</sub>, at 0°, or in Et<sub>2</sub>O, Me<sub>2</sub>CO or EtOAc. [See Gunghaly et al. *Arch Biochem Biophys* **38** 75 1952, *Beilstein* **6** IV 4133.]

**Vitamin A aldehyde [all-trans-retinal; 3,7-dimethyl-9-(2,6,6-trimethyl-1-cyclohexenyl)-2,4,6,8-nonatetraen-1-al]** [116-31-4] **M 284.4, m 61-64°**. The aldehyde is separated from retinol by column chromatography on water-deactivated alumina. Elute with 1-2% acetone in hexane, or on TLC plates of silica gel G and using the same eluting solvent. It crystallises from petroleum ether or *n*-hexane as yellow-orange crystals, and the UV in hexane has  $\lambda_{\max}$  at 373nm ( $A_{1cm}^{1\%}$  1548) and 368nm ( $\epsilon$  48,000). It is an **irritant** and is light sensitive. Store it in sealed ampoules under N<sub>2</sub>. The *semicarbazone* forms yellow crystals from CHCl<sub>3</sub>/Et<sub>2</sub>O or EtOH, **m 199-201°(dec)**. The *9-cis-isomer* [514-85-2] and the *13-cis-isomer* [472-86-6] [ $\lambda_{\max}$  at 375nm ( $\epsilon$  1,250) in EtOH] are also available commercially. [*Beilstein* **7** III 1742.]

**Vitamin A palmitate (retinyl palmitate)** [79-81-2] **M 524.9, m 28-29°,  $\epsilon_{1cm}^{1\%}$  (all-trans) 1000 (325 nm) in EtOH**. The palmitate is separated from retinol by column chromatography on water-deactivated alumina with hexane containing a very small percentage of acetone. It is also chromatographed on TLC silica gel G, using petroleum ether/isopropyl ether/acetic acid/water (180:20:2:5) or petroleum ether/acetonitrile/acetic acid/water (190:10:1:15) to develop the chromatogram. It is then recrystallised from propylene at low temperature (below -47°). [*Beilstein* **6** IV 4135.]

**Vitamin B<sub>1</sub> Hydrochloride [Aneurine hydrochloride, Thiamine hydrochloride, 3{(4-amino-2-methyl-5-pyrimidinyl)methyl}-4-methylthiazolium chloride monohydrochloride]** [67-03-8] **M 337.3, m 248°(dec), 249-250°, monohydrate m 135°(dec), pK<sub>1</sub><sup>25</sup> 4.8, pK<sub>2</sub><sup>25</sup> 9.2**. The hydrochloride crystallises from 95% EtOH (solubility is *ca* 1%). The *monohydrate* is dehydrated at 100° *in vacuo* over H<sub>2</sub>SO<sub>4</sub>, but is *hygroscopic* and picks up one molecule of H<sub>2</sub>O readily. It can be sterilised at 100° if the pH of the solution is below 5.5. The *nitrate* has **m 196-200°(dec)** and is more stable than the hydrochloride. The *picrolonate* crystallises from H<sub>2</sub>O and is *dimorphic*, **m 164-165° and 228-229°(dec)**. [Todd & Bergel *J Chem Soc* 364, 367 1937, *J Am Chem Soc* **58** 1063, 1504 1936, **59** 526 1937, *Beilstein* **27** IV 1766.]

**Vitamin B<sub>2</sub>** [Riboflavin, Lactoflavin, 6,7-dimethyl-9-(D-1'-ribyl)isoalloxazine] [83-88-5] **M 376.4, m 278-282°**(dec with darkening at 240°), 281-282°,  $[\alpha]_{\text{D}}^{25} -9.8^{\circ}$  (H<sub>2</sub>O),  $[\alpha]_{\text{D}}^{25} -112^{\circ}$  to  $-125^{\circ}$  (c 2.5, 0.02N NaOH),  $[\alpha]_{\text{D}}^{20} -59^{\circ}$  (c 0.23, AcOH), **pK<sub>1</sub> 1.7, pK<sub>2</sub> 9.69 (10.2, acidic NH)**. It crystallises from H<sub>2</sub>O as a yellow-orange powder in three different forms with differing amounts of H<sub>2</sub>O. It melts if placed in an oil bath at 250°, but decomposes at 280° if heated at a rate of 5°/minute. It is also purified by crystallisation from 2M acetic acid, then extracted with CHCl<sub>3</sub> to remove lumichrome impurity. [Smith & Metzler *J Am Chem Soc* **85** 3285 1963.] Its solubility in H<sub>2</sub>O is 1g in 3-15L depending on the crystal structure. Its solubility in EtOH at 25° is 4.5mg in 100ml. Store it in the dark because it is decomposed to lumichrome by UV light. [Pearson *The Vitamins* **vol V** pp1-96 1967 and **vol VII** pp 1-96 1972, Gyögy and Pearson eds, Academic Press, *Beilstein* **26** IV 2542.]

**Vitamin B<sub>3</sub> (vitamin PP, Niacin)**. See nicotinamide in "Heterocyclic Compounds", Chapter 4.

**Vitamin B<sub>6</sub> hydrochloride** (adermine, pyridoxine HCl, 3-hydroxy-4,5-bis[hydroxymethyl]-2-methylpyridine HCl) [58-56-0] **M 205.6, m 208-208.5°, 208-209°(dec), 209-210°(dec), 205-212° (sublimes), pK<sub>1</sub><sup>25</sup> 5.0 (3-OH), pK<sub>2</sub><sup>25</sup> 8.96 (pyridinium<sup>+</sup>)**. Purify the vitamin by recrystallisation from EtOH/Me<sub>2</sub>CO, *n*-BuOH or MeOH/Et<sub>2</sub>O. Its solubility in H<sub>2</sub>O is 22%, and in EtOH it is 1.1%. It is insoluble in Et<sub>2</sub>O and CHCl<sub>3</sub>. Acidic aqueous solutions are stable at 120°/30minute. The *free base* has **m 159-160°** after recrystallisation from Me<sub>2</sub>CO and sublimation at 140-145°/0.0001mm. It has UV with  $\lambda_{\text{max}}$  at 290nm ( $\epsilon$  84,000), 253 and 325nm ( $\epsilon$  3700 and 7100) in 0.1N aqueous HCl. [Khua & Wendt *Chem Ber* **71** 780 1938, **72** 311 1939, Harris & Folkers *J Am Chem Soc* **61** 1242 1939, Harris et al. *J Am Chem Soc* **62** 3198 1940, *Beilstein* **21/5** V 492.] See also **Pyridoxal-5'-phosphate H<sub>2</sub>O** above.

**Vitamin B<sub>12</sub> (cyanocobalamine,  $\alpha$ -[5,6-dimethylbenzimidazolyl]cyano cobamide)** [68-19-9] **M 1355.4, m darkens at 210-220° and does not melt below 300°,  $[\alpha]_{\text{D}}^{23} -59^{\circ}$  (H<sub>2</sub>O)**. Vitamin B<sub>12</sub> crystallises from de-ionized H<sub>2</sub>O, with a solubility in H<sub>2</sub>O of 1g/80g, and is dried under vacuum over Mg(ClO<sub>4</sub>)<sub>2</sub>. The dry red crystals are *hygroscopic* and can absorb ~12% of H<sub>2</sub>O. A solution at pH 4.5-5 can be autoclaved for 20 minutes at 120° without decomposition. Aqueous solutions are stabilised by addition of (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>. [Golding *Comprehensive Organic Chem* Vol 5 (Ed. Haslam; Pergamon Press, NY, 1979) pp 549-584.] *Alternatively*, an aqueous solution of the coenzyme can be concentrated, if necessary in a vacuum at 25° or less, until the concentration is 0.005 to 0.01M (as estimated by the OD at 522nm of an aliquot diluted with 0.01M K-phosphate buffer pH 7.0). If crystals begin to form on the walls of the container, they should be re-dissolved with a little H<sub>2</sub>O. The concentrated solution is placed in a glass stoppered flask and diluted with 5 volumes of Me<sub>2</sub>CO. After 2-3 hours at 3° it is centrifuged (10,000xg/10minutes) in Me<sub>2</sub>CO-insoluble plastic tubes to remove some amorphous precipitate. The clear supernatant is inoculated with a small crystal of the vitamin and allowed to crystallise overnight at 3°. Crystals are formed on the walls and the bottom of the container. A further 2 volumes of Me<sub>2</sub>CO are added and set aside at 3° to further crystallise. Crystallisation is followed by observing the OD<sub>522</sub> of the supernatant. When the OD falls to 0.27, then *ca* 94% of the crystals have separated. The supernatant is decanted (saved for obtaining a second crop), and the crystals are washed with a little cold 90% aqueous Me<sub>2</sub>CO (2x), 100% Me<sub>2</sub>CO (2x), Et<sub>2</sub>O (2x) at which time the crystals separate from the glass walls. Allow them to settle and remove residual Et<sub>2</sub>O with a stream of dry N<sub>2</sub>. The process can be repeated if necessary. The crystals can be dried in air or in a vacuum for 2 hours over silica gel at 100° with an 8-9% weight loss. [Barker et al. *Biochemical Preparations* **10** 33 1963.] This material gives a single spot on paper chromatography [see Weissbach et al. *J Biol Chem* **235** 1462 1960.] The vitamin is soluble in H<sub>2</sub>O (16.4mM at 24°, 6.4mM at 1°), in EtOH and PhOH but insoluble in Me<sub>2</sub>CO, Et<sub>2</sub>O, CH<sub>2</sub>Cl<sub>2</sub> and dioxane. Its UV has  $\lambda_{\text{max}}$  at 260, 375 and 522nm ( $\epsilon$  34.7 x 10<sup>6</sup>, 10.9 x 10<sup>6</sup> and 8.0 x 10<sup>6</sup>/mole) in H<sub>2</sub>O. The dry crystals are stable for months in the dark, but aqueous solutions decompose on exposure to VIS or UV light or alkaline CN<sup>-</sup>, but are stable in the dark at pH 6-7. The vitamin is inactivated by strong acids or alkalis. [Barker et al. *J Biol Chem* **235** 480 1960, see also *Vitamin B<sub>12</sub>* (Zagalak & Friedrich Eds) Walter de Gruyter, Berlin 1979, *Beilstein* **26** IV 3117.]

**Vitamin C**. See ascorbic acid entry in "Carbohydrates" in this chapter.

**Vitamin D<sub>2</sub>** [50-14-6] **M 396.7, m 114-116°,  $[\alpha]_{\text{D}}^{20} +122^{\circ}$  (c 4, EtOH)**. It is converted into its 3,5-dinitrobenzoyl ester and crystallised repeatedly from acetone. The ester is then saponified and the free vitamin is

isolated. [Laughland & Phillips *Anal Chem* **28** 817 1956, *Beilstein* **6** IV 4404.]

**Vitamin D<sub>3</sub>** (cholecalciferol, Calcitol, activated 7-dehydrocholesterol, (+)-vitamin D<sub>3</sub>) [67-97-0] **M 384.6, m 83-85°**,  $[\alpha]_{546}^{20} +126^{\circ}$  (c 2, EtOH). It is converted into its 3,5-dinitrobenzoyl ester and crystallised repeatedly from acetone. The ester is then saponified and the free vitamin is isolated. Store it in sealed ampoules under argon below 8°. It acts through a receptor which modulates differentiation and proliferation of normal and neoplastic cells. [Laughland & Phillips *Anal Chem* **28** 817 1956, DeLuca & Schones *Ann Rev Biochem* **52** 411 1983, *Beilstein* **6** III 2811, **6** IV 4149.]

**Vitamin E** (2*R*,4'*R*,8'*R*- $\alpha$ -tocopherol, natural active isomer) [59-02-9] **M 430.7, m 2.5-3.5°, b 200-220°/0.1mm, 200°/0.005mm, d<sub>4</sub><sup>25</sup> 0.950, n<sub>D</sub><sup>25</sup> 1.5045,  $[\alpha]_{D}^{25} +3.58^{\circ}$  (c 1, \*C<sub>6</sub>H<sub>6</sub>)**. Vitamin E is a viscous yellow oil which is distilled at high vacuum. It has  $\lambda_{max}$  at 294nm ( $E_{1cm}^{1\%}$  71). It is oxygen and light sensitive and is best stored as its stable *D*- $\alpha$ -acetate [58-95-7] which is purified by evaporative distillation at **b** 180-200°(bath temperature)/0.7mm, and has  $[\alpha]_{D}^{25} +3.3^{\circ}$  (c 5.1, EtOH). It forms needles at -30° and has **m** 26.5-27.5°,  $[\alpha]_{D}^{25} +0.25^{\circ}$  (c 10, CHCl<sub>3</sub>). [NMR: Cohen et al. *Helv Chim Acta* **64** 1158 1981, Burton & Ingold *Acc Chem Res* **19** 194 1986, Karrer et al. *Helv Chim Acta* **21** 520 1938, Robeson *J Am Chem Soc*, **64** 1487 1942, **65** 1660 1943.] Of the eight isomers the *D*- $\alpha$ -isomer is the most active. [See W. Friedrich "Vitamins" Walter de Guyter Publ, Berlin 1988.] [*Beilstein* **17/4** V 168.]

**DL-Vitamin E acetate** (DL- $\alpha$ -tocopheryl acetate) [7695-91-2] **M 472.8, m -27.5°, b 194-196°/0.01mm, 222-224°/0.3mm, d<sub>4</sub><sup>20</sup> 0.958, n<sub>D</sub><sup>20</sup> 1.4958**. It is a viscous liquid which is purified by distillation under high vacuum under N<sub>2</sub> or argon and stored in sealed ampoules in the dark. It is considerably more stable to light and air than the parent unacetylated vitamin. It is insoluble in H<sub>2</sub>O but freely soluble in organic solvents. All eight stereoisomers have been synthesised. The commercially pure *d*- $\alpha$ -tocopheryl acetate (2*R*,4'*R*,8'*R*) has **b** 180-200°/0.7mm and  $[\alpha]_{D}^{20} +3.9^{\circ}$  (c 5, EtOH); see above. [Cohen et al. *Helv Chim Acta* **64** 1158 1981, *Beilstein* **17/4** V 169.]

**Vitamin K<sub>1</sub>** (2-methyl-3-phytyl-1,4-naphthoquinone) [84-80-0] **M 450.7, m -20°, b 141-140/0.001mm, b 140-145°/10<sup>-3</sup> mm, d<sub>4</sub><sup>25</sup> 0.967, n<sub>D</sub><sup>25</sup> 1.527,  $[\alpha]_{D}^{20} -0.4^{\circ}$  (c 57.5, \*C<sub>6</sub>H<sub>6</sub>)**. Vitamin K<sub>1</sub> is a yellow viscous oil, which can be distilled at high vacuum practically unchanged. It is insoluble in H<sub>2</sub>O, but soluble in common organic solvents. Store it in the dark under N<sub>2</sub> as it is oxygen sensitive. It has  $A_{1cm}^{1\%}$  328 at 248nm. [Fieser et al. *J Am Chem Soc* **61** 2557 1939, Hirschmann et al. *J Am Chem Soc* **76** 4592 1954, Isler & Doebel *Helv Chim Acta* **27** 225 1954, *Beilstein* **7** IV 2496.]

**Vitamin K<sub>3</sub>** (2-methyl-1,4-naphthoquinone, Menadione, Menaphthone) [58-27-5] **M 172.2, m 105-106°, 105-107°**. Recrystallise it from 95% EtOH, or MeOH after filtration. It forms bright yellow crystals which are decomposed by light. Its solubility in EtOH is 1.7% and in \*C<sub>6</sub>H<sub>6</sub> it is 10%. It **IRRITATES** mucous membranes and skin. [Fieser *J Biol Chem* **133** 391 1940, *Beilstein* **7** IV 2430.]

**Vitamin K<sub>5</sub>** (4-Amino-2-methyl-1-naphthol hydrochloride) [130-24-5] **M 209.6, m 283°(dec), pK<sub>Est(1)</sub> ~5.6 (NH<sub>2</sub>), pK<sub>Est(2)</sub> ~10.4 (OH)**. Crystallise it from dilute HCl. [Sah *Rec Trav Chim Pays Bas* **59** 458 1941, Sah *Rec Trav Chim Pays Bas* **60** 373 1940, Veldstra & Wiardi *Rec Trav Chim Pays Bas* **61** 547 1942, Veldstra & Wiardi *Rec Trav Chim Pays Bas* **62** 75 1943, *Beilstein* **13** III 1921.]

**Vitamin P** [(+)-Rutin, quercetin-3-rubinoside] [153-18-4] **M 610.5, 664.5 (trihydrate), m 188-189°, 195-196°,  $[\alpha]_{546}^{20} +13^{\circ}$  (c 5, EtOH),  $[\alpha]_{D}^{24} -37.6^{\circ}$  (c 1.24, pyridine), (polyphenolic flavone pKs 7—10)**. The vitamin crystallises from MeOH or water/EtOH, dry it in air, then dry it further for several hours at 110° or in high vacuum at 120°. It forms yellow crystals from EtOH/Me<sub>2</sub>CO/H<sub>2</sub>O (2:1:1). It has also been purified by passing (0.5g) through a Kieselgel column (30 x 5cm) with EtOAc/MeOH/H<sub>2</sub>O (100:20:15), and after 750ml have passed through, the yellow fraction of 250ml gives the glycoside (0.3g) on evaporation. [Hörhammer et al. *Chem Ber* **101** 1183 1968, Marini-Bettòlo *Gazz Chim Ital* **80** 631 1950, *Beilstein* **18/5** V 519.]

**Vitamin U** (S-Methyl-L-methionine chloride) [*S*- 1115-84-0, *S*-ion- 6708-35-6, *RS*- 44901-24-8] **M 199.5, m 135° (dec), 139° (dec),  $[\alpha]_{D}^{23} +33^{\circ}$  (0.2M HCl), pK<sub>1</sub> 1.9, pK<sub>2</sub> 7.9**. Likely impurities are methionine, methionine

sulfoxide and methionine sulfone. Crystallise Vitamin U from water by adding a large excess of EtOH. It is *hygroscopic*. It should be stored in a cool, dry place and protected from light. It has also been purified on a column of Dowex 50 H<sup>+</sup> form, washed with H<sub>2</sub>O and eluted with 6N ammonia, lyophilised, the residue is dissolved in dilute HCl, lyophilised again and then it is obtained as colourless prisms by vapour diffusion of Et<sub>2</sub>O in a 1:1 (v/v) mixture of MeOH/EtOH in which the *hydrochloride* is dissolved [Del Re *Acta Cryst B* **33** 3289 1977]. The *iodide salt*, **m** ~150°(dec), is obtained by dissolving it in 50% aqueous EtOH and adding ~3.5 volumes of absolute EtOH. The *bromide salt* has **m** 140°(dec) [Toennies & Kolb *J Amer Chem Soc* **67** 849 1945].

***dl*-Warfarin (Coumafene, 4-hydroxy-3-(3-oxo-1-phenylbutyl)-2H-1-benzopyran-2-one)** [81-81-2] **M 308.3, m 162-164°, pK<sup>20</sup> 4.85 (the 4-mercapto derivative has pK<sup>20</sup> 6.60).** *dl*-Warfarin crystallises from EtOH or MeOH. Its UV has  $\lambda_{\max}$  at 308nm ( $\epsilon$  13,610) in H<sub>2</sub>O. The *acetate* has **m** 117-118°, the *O-triflate* has **m** 90-91°, and the *2,4-dinitrophenylhydrazone* has **m** 215-216°. It is an effective anticoagulant and rodenticide. [West et al. *J Am Chem Soc* **83** 2676 1961, HPLC: Banfield & Rowland *J Pharm Sci* **72** 921 1983, *Beilstein* **17** III/IV 6794.]

***R*(+)-Warfarin** has **m** 170-171°,  $[\alpha]_{\text{D}}^{25}$  +149° (c 1.2, in 0.5N NaOH), +25.5° (c 2, AcOH), +15.5° (c 2, MeCN) and -19° (c 1.1, propan-2-ol), and ***S*(-)-Warfarin (more active enantiomer)** has **m** 172-173°,  $[\alpha]_{\text{D}}^{25}$  -148° (c 1.2, in 0.5N NaOH), -25.5° (c 2, AcOH), -15.5° (c 2, MeCN) and +19° (c 1.1, propan-2-ol). *dl*-Warfarin is resolved *via* recrystallisation of the quinidine salt, and the free acids are recrystallised (70g) from 600ml of 80% aqueous Me<sub>2</sub>CO. Large prismatic crystals of the pure enantiomers are obtained by slow crystallisation from Me<sub>2</sub>CO or AcOH. The solubilities of the pure enantiomers at 25° are 11.2% in Me<sub>2</sub>CO and 2.6% in AcOH, whereas the racemate has solubilities of 6.5% in Me<sub>2</sub>CO and 2% in AcOH. The IR spectra are the same with  $\nu_{\max}$  (CHCl<sub>3</sub>) at 2.78 (w), 5.88, 6.16 and 6.38 $\mu$ . [West et al. *J Am Chem Soc* **83** 2676 1961, Cbz-proline diastereoisomeric esters were used for HPLC analysis: Banfield & Rowland *J Pharm Sci* **72** 921 1983.] **Poisonous, anticoagulant and rodenticide.**

**Xanthopterin monohydrate (2-amino-4,6-dihydroxypteridine, 2-aminopteridin-4,6(1*H*,5*H*)-dione)** [5979-01-1 (H<sub>2</sub>O), 119-48-8 (anhydrous)] **M 197.2, m <300°, pK<sub>1</sub> 1.6 (basic), pK<sub>2</sub> 6.59 (acidic), pK<sub>3</sub> 9.31 (acidic)(anhydrous species), and pK<sub>1</sub> 1.6 (basic), pK<sub>2</sub> 8.65 (acidic), pK<sub>3</sub> 9.99 (acidic)(7,8-hydrated species).** Purification is as for isoxanthopterin (see [529-69-1] "Heterocyclic Compounds", Chapter 4). It is crystallised by acidifying a hot ammoniacal solution with formic acid, and collecting the crystals by centrifugation followed by washing with EtOH, ether and drying at 100° *in vacuo*. Its R<sub>F</sub> values on paper chromatography are 0.15 (*n*-PrOH, 1% aqueous NH<sub>3</sub>, 2:1), 0.36 (*n*-BuOH/AcOH/H<sub>2</sub>O, 4:1:1) and 0.47 (3% aqueous NH<sub>3</sub>). [Inoue & Perrin *J Chem Soc* 260 1962, Inoue *Tetrahedron* **20** 243 1964, see also Blakley *Biochemistry of Folic Acid and Related Pteridines* North Holland Publ Co, Amsterdam 1969, *Beilstein* **26** II 313, **26** III/IV 4000.]

**Xanthotoxin (Methoxalen, 9-methoxypsoralen, Ammoidin, 9-methoxyfuro[3,2-*g*][1]benzopyran-7-one)** [298-81-7] **M 216.2, m 146-148°, 148°, 148-149°.** Purify xanthotoxin by recrystallisation from \*C<sub>6</sub>H<sub>6</sub>/petroleum ether (b 60-80°) to give silky needles, or from EtOH/Et<sub>2</sub>O to give rhombic prisms or from hot H<sub>2</sub>O to give needles. It is soluble in aqueous alkali due to ring opening of the cyclic lactone but recyclises upon acidification. It has UV with  $\lambda_{\max}$  (EtOH) at 219, 249 and 300nm (log  $\epsilon$  4.32, 4.35 and 4.06) and <sup>1</sup>H NMR in CDCl<sub>3</sub> with  $\delta$  at 7.76 (d, 1H, *J* = 10Hz), 7.71 (d, 1H, *J* = 2.5Hz), 7.38 (s, 1H), 6.84 (d, 1H, *J* = 2.5Hz), 6.39 (d, 1H, *J* = 10Hz) and 4.28 (s, 3H). [Nore & Honkanen *J Heterocycl Chem* **17** 985 1980.] It is a DNA intercalator, is used in the treatment of dermal diseases, and is a **human carcinogen** [Tessman et al. *Biochemistry* **24** 1669 1985]. [*Beilstein* **19** I 711, **19/6** V 15.]

**Zeatin (trans-*N*<sup>6</sup>-[4-hydroxy-3-methylbut-2-en-1-yl]adenine)** [1637-39-4] **M 219.3, m 207-208, 209-209.5°, pK<sub>1</sub> 4.4 (basic), pK<sub>2</sub> 9.8 (acidic).** Purify zeatin by recrystallisation from EtOH or H<sub>2</sub>O. Its UV has  $\lambda_{\max}$  at 207 and 275nm ( $\epsilon$  1400 and 14,650) in 0.1N aqueous HCl; 212 and 270nm ( $\epsilon$  17,050 and 16,150) in aqueous buffer pH 7.2; and 220 and 276nm ( $\epsilon$  15,900 and 14,650) in 0.1N aqueous NaOH. The *picrate* has **m** 192-194° (from H<sub>2</sub>O) from which zeatin can be recovered by treatment with Dowex-I x 8 (200-400 mesh, OH<sup>-</sup> form).

[Letham et al. *Aust J Chem* **22** 205 1969, *Proc Chem Soc (London)* 230 1964, Shaw & Wilson *Proc Chem Soc (London)* 231 1964.] It is a cell division factor (plant growth regulator) [Letham & Palni *Ann Rev Plant Physiol* **34** 163 1983] and inhibits mitochondrial function [Miller *Plant Physiol* **69** 1274 1982]. The commercially available *trans*-9-riboside derivative, [6025-53-2], **M 351.4**, is a cytokine which separates from aqueous AcOH with **m** 177-179°. Its solubility in AcOH is ~5%. [McDonald & Morris *Methods Enzymol* **100** 347 1985.]

## CHAPTER 8

### NANOMATERIALS AND NANOTECHNOLOGY

#### INTRODUCTION

*Nanomaterials* are substances that are, or have been, reduced in size to the range from 1nm to ~100nm (i.e. 1 to ~100 nanometers, or 1 to ~100 x 10<sup>-9</sup> meters). *Nanotechnology* is the science and applications of nanomaterials, and is growing at an ever increasing pace. At this particle size the properties of materials can be altered dramatically. Properties such as solubility, reactivity, spectroscopy, electrical and magnetic, transport through membranes etc. are generally different from those of the same materials with large particle size. The applications of materials of nano size have escalated in the last fifteen or so years and are currently gaining momentum. The technology has broad applications in performance materials, health, consumer products, water, information technology and energy. The discovery of *fullerenes* (1996 Nobel Prize for Chemistry to H.W. Kroto, R.E. Smalley and R.F. Curl) and of *Graphene* (2010 Nobel Prize for Physics to A. Geim and K. Novoselov) are very important factors in the development of nanotechnology. Not very many nanomaterials have so far been included in chemical catalogues of commercially available substances, e.g. as compared with catalysts (although some catalysts are now available as nanoparticles), hence the relatively smaller size of this chapter. Fullerenes and related materials such as carbon nanotubes have found applications in medicine for the transport of drugs and biological materials to targeted sites and some readily available derivatives are described in this chapter, together with other commercially available nanoparticulate substances, e.g. catalysts, magnetic and electrically conducting thin films. General safety aspects of handling nanomaterials have been briefly addressed.

#### IDENTIFICATION AND MEASUREMENT OF NANOMATERIALS

The 1 to ~ 100nm (10 to ~1000Å) particles are smaller than the wavelength of UV light, and thus the use of appropriate small particles e.g. electrons, or radiation e.g. X-rays is necessary to identify and measure these particles. Instrumentation for these purposes has been developed and the resolutions are continuously being improved. These instruments can measure very accurately, and depending on the physical principles of the machines, they measure different parameters of the nanoparticles, e.g. diameter (if spherical), length (if rod like), width (if bulky) etc, and not necessarily all parameters at the same time with one instrument. The commonly used instruments include the *Transmission Electron Microscope* (TEM) in which the electron beam runs past the particles placed on a transparent surface, and then through the surface giving a 2D image and magnifying one million times. In the *Scanning Electron Microscope* (SEM) and *Scanning Tunneling Microscope* (STM, or *Spectroscopy* STS) the electron beam scans the surface of the particles to provide a 3D picture, e.g. of hairs on a fly. *Dynamic Light Scattering* (DLS) uses light scattering through solutions of nanoparticles. Another technique, *Scattering Probe Microscopy* (SPM), allows scanning with a probe over a surface and includes *Atomic Force Microscopy* (AFM, or *non-contact atomic force microscopy* NC-AFM) where a ceramic or semiconductor tip scans the surface (much like a phonograph needle scans a record), and measures the atomic forces beneath it. The NC-AFM tip is attached to a very small tuning fork that oscillates at amplitudes of 0.2 to 0.5Å, smaller than the diameter of an atom. When the tip is brought very close to a molecule on a surface, this frequency alters slightly thus *detuning* due to the atomic forces between the tip and the molecule to give a complete picture of the molecule on the surface. By placing a CO (carbon monoxide) molecule at the tip, and keeping it there throughout the imaging process, very high resolution can be obtained. Gross and coworkers [Gross et al. *Science* **325** 1110 2009] obtained the complete image of pentacene on a copper surface in this way, and succeeded in imaging the amount of charge on a single atom [Gross et al. *Science* **324** 1428 2011]. The sensitivity of *Magnetic Resonance Imaging* (MRI) has been considerably increased (with a resolution of ~4nm), by Rugar and coworkers [Degen et al. *Proc Natl Acad Sci* **106** 1313 2009, cf H. Birch *MRI on the Nanoscale*, RSC Chemistry World **6**(2) 28 2009, J. M. Crow *Picture perfect pentacene*,

RSC *Chemistry World* 6(1) 462011] who have developed the *Magnetic Resonance Force Microscopy* (MRFM) technique. Here the sample is placed at the end of a tiny silicon cantilever to which is applied, *via* a magnetic tip, a radio frequency magnetic field. The spin of the nuclei alternate (*up* and *down*) when the frequency reaches their ‘wobbling’ frequencies making the cantilever vibrate. Photon density 3D images of the specimen (e.g. tobacco mosaic virus, *ca* 18 nanometers across) can be obtained by using a laser to detect the cantilever vibrations. The technique does not affect the sample and gives superior images to those from STM and AFM. [A.I. Kirkland and J. I. Hutchinson (eds) *Nanocharacterisation* RSC Publishing UK, 2007, ISBN 9781847557926, 1847557929; A. Turley *Sizing it up*, RSC *Chemistry World* 8(3) 50 2011; M. Thompson RSC *The Characterisation of Nanoparticles; Analytical Methods Committee Technical Briefs* 48, December 2010; ISSN 1757-5958, <http://bit.ly/AMCparticle>). Other techniques include *Near-Infrared Spectroscopy* (NIR) and solution phase NIR [Iktis et al. *Nano Lett* 3 309 2003], general *Electron Microscopy* (EM), *Field Emission Microscopy* (FEM), *FT-Raman Scattering Spectroscopy*, *Nuclear (proton) Magnetic Resonance Spectroscopy* (NMR for derivatised carbon nanotubes, CNTs), *Thermo Gravimetric Analysis* (TGA), *Energy Dispersive X-ray Spectroscopy* (EDX), as well as other optical techniques (see Bibliography at the end of this chapter).

Preparation of the specimen is very important since the atoms or/and molecules have to be spread over the surface in as close to one atom or molecule thick. Where thicknesses (*t*) of 7 to ~13nm need to be measured *Grazing Incident X-Ray Reflectometry* (GICRR/XRR) provides an initial measure of particle lengths (*l*) of 1 to 35nm. *Spectroscopic Ellipsometry* (SE, *t* = ~ 10.5 to 12.3nm, *l* = 1 to 10.5nm), *Scanning Ion Mass Spectrometry* (SIMS, *t* ~ 10.4 to 10.8nm, *l* = ~12 to 13 nm), *Transmission Electron Microscopy* (TEM, *t* ~ 7.5 to 12.5nm, *l* = 12.5 to 18.5 nm), *X-Ray Fluorescence Spectroscopy* (XFS, *t* ~ 7.5 to 9.3nm, *l* = 16.0 to 19.5 nm), *X-Ray Photoelectron Spectroscopy* (XPS, *t* ~ 9.2 to 11.7, *l* = 18.5 to 25.5 nm) and higher resolution *Grazing Incident X-Ray Reflectometry* (GIXRR/XRR, *t* ~ 9.2 to 11.7, *l* = 24.0 to 36.0 nm) provide the approximate *t* and *l* values stated in the brackets. It must be noted that these values are approximate as they do vary depending on the nature of the materials that are measured.

**Average particle size:** The accepted method for determining the primary particle size of nanomaterials is by TEM analysis. However, this is expensive and time-consuming. Studies have shown that an average primary particle size for spherical or cubic shaped dry nanoparticles of roughly uniform shape and size can be calculated from the equation:

$$\text{Average particle size (nanometers)} = (6/\text{BET surface area}^* \times \text{density}^{**}) \times 1000$$

Where surface area\* is in sq. m/g, and density\*\* is in g/cc. *Micromeritics Analytical Services* (MAS) will measure BET surface area by a gas absorption procedure and the real density by gas pycnometry. This gives a good estimate of average primary particle size. For further information see [www.micromeritics.com](http://www.micromeritics.com), Micromeritics Analytical Services Downloads.

## FULLERENES AND RELATED SUBSTANCES

Fullerenes and related substances are made by heating carbon (e.g. soot, graphite) to form nano-sized carbon ‘balls’. These nanoparticles look like ‘footballs’ in being constructed of five (pentagon) and/or six (hexagon) carbon rings. Their carbon ring sizes make for stable ‘ball-like’ molecules of 60 or more carbon atoms and have some solubility in organic solvents. These carbon allotropes have a variety of exotic properties including electrochemical [Echegoyen & Echegoyen *Acc Chem Res* 31 593 1998], ferromagnetism [Allemand et al. *Science* 253 301 1991], superconductivity [Hebard et al. *Nature* 350 600 1991], photophysical [Guldi *Chem Commun* 321 2000], anti-HIV bioactivity [Friedman et al. *J Am Chem Soc* 115 6506 1993], and their ‘aromatic’ nature allows them to react accordingly and be functionalised (see below) for further chemical reactions [Kordatos et al. *J Org Chem* 66 4915 2001]. New and useful applications are continually being found in nano science, medicine, pharmaceuticals, and nanotechnology [see references below].

### FULLERENES

**Fullerene C<sub>60</sub>** (Buckminsterfullerene C<sub>60</sub>, Footballene, Buckyball 60) [99685-96-8] M 720.66 and **Fullerene C<sub>70</sub>** [115383-22-7] M 840.77. **Purification procedures: (a)** This was purified from the soluble toluene extract (400mg) of the soot (Fullerite), formed from resistive heating of graphite, by adsorption on neutral alumina (100g, Brockmann I, 60 x 8cm). Elution with toluene/hexane (5:95 v/v) gives *ca* 250mg of quite pure C<sub>60</sub>. It has characteristic spectral properties (see below). Further elution with toluene/hexane (20:80

v/v, i.e. increased polarity of solvent) provides 50mg of "pure" C<sub>70</sub> [Allemand et al. *J Am Chem Soc* **113** 1050 1991].

Chromatography on alumina can be improved by using conditions which favour adsorption rather than crystallisation. Thus the residue from toluene extraction (1g) in CS<sub>2</sub> (ca 300ml) is adsorbed on alumina (375g, standard grade, neutral ca 150 mesh, Brockmann I) and loaded as a slurry in toluene/hexanes (5:95 v/v) to a 50 x 8cm column of alumina (1.5Kg) in the same solvent. To avoid crystallisation of the fullerenes, 10% of toluene in hexane is added quickly followed by 5% of toluene in hexane after the fullerenes had left the loading fraction (2-3 hours). With a flow rate of 15ml/minute the purple C<sub>60</sub> fraction is eluted during a 3-4 hour period. Evaporation of the eluates gives 550-630mg of product which, after recrystallisation from CS<sub>2</sub>/cyclohexane yields 520-600mg of C<sub>60</sub> which contains adsorbed solvent. On drying at 275<sup>o</sup>/10<sup>-3</sup>mm for 48 hours a 2% weight loss is observed although the C<sub>60</sub> still contains traces of solvent. Further elution of the column with 20% of toluene in hexane provides 130mg of C<sub>70</sub> containing 10-14% of C<sub>60</sub> (by <sup>13</sup>C NMR). This was re-chromatographed as above using a half scale column and adsorbing the 130mg in CS<sub>2</sub> (20ml) on alumina (24g) which gave 105mg of recrystallised C<sub>70</sub> (containing 2% of C<sub>60</sub>). The purity of C<sub>60</sub> can be improved further by washing the crystalline product with Et<sub>2</sub>O and Me<sub>2</sub>CO followed by recrystallisation from \*C<sub>6</sub>H<sub>6</sub> and vacuum drying at high temperatures.

(b) Carbon soot from resistive heating of a carbon rod in a partial helium atmosphere (0.3bar) under specified conditions is extracted with boiling \*C<sub>6</sub>H<sub>6</sub> or toluene, filtered and the red-brown solution is evaporated to give crystalline material in 14% yield which is mainly a mixture of fullerenes C<sub>60</sub> and C<sub>70</sub>. Chromatographic filtration of the 'crude' mixture with \*C<sub>6</sub>H<sub>6</sub> allows no separation of components, but some separation was observed on silica gel TLC with *n*-hexane or *n*-pentane, but not cyclohexane as eluents. Analytical HPLC with hexanes (5µm Econosphere silica) gave satisfactory separation of C<sub>60</sub> and C<sub>70</sub> (retention times of 6.64 and 6.93 minutes respectively) at a flow rate of 0.5ml/minute and using a detector at 256nm. HPLC indicated the presence of minor (<1.5% of total mass) unidentified C<sub>n</sub> species with retention times of 5.86 and 8.31 minutes. Column chromatography on flash silica gel with hexane gives a few fractions of C<sub>60</sub> with ≥95% purity, but later fractions contain mixtures of C<sub>60</sub> and C<sub>70</sub>. These can be obtained in 99.85 and >99% purity, respectively, by column chromatography on neutral alumina. [Ajie et al. *J Phys Chem* **94** 8630 1990.]

(c) Separation of C<sub>60</sub> and C<sub>70</sub> can be achieved by HPLC on a dinitroanilinopropyl (DNAP) silica (5µm pore size, 300Å pore diameter) column with a gradient from *n*-hexane to 50% CH<sub>2</sub>Cl<sub>2</sub> using a diode array detector at wavelengths 330nm (for C<sub>60</sub>) and 384nm (for C<sub>70</sub>). [Cox et al. *J Am Chem Soc* **113**, 2940, 1991.]

Soxhlet extraction of the "soot" is a good preliminary procedure, or if material of only ca 98% purity is required. Soxhlet extraction with toluene is run (20 minutes per cycle) until colourless solvent fills the upper part of the Soxhlet equipment (10 hours). One-third of the toluene remains in the pot. After cooling, the solution is filtered through a glass frit. This solid (purple in toluene) is ca 98% C<sub>60</sub>. This powder is again extracted in a Soxhlet using identical conditions as before, and the C<sub>60</sub> is recrystallised from toluene to give 99.5% pure C<sub>60</sub>. C<sub>70</sub> has greater affinity than C<sub>60</sub> for toluene. [Coustel et al. *J Chem Soc, Chem Commun* 1402 1992.]

(d) Purification of C<sub>60</sub> from a C<sub>60</sub>/C<sub>70</sub> mixture was also achieved by dissolving it in an aqueous solution of γ (but not β) cyclodextrin (0.02M) with refluxing. The rate of dissolution (as can be followed by UV spectroscopy) is quite slow and constant up to 10<sup>-5</sup>M of C<sub>60</sub>. The highest concentration of C<sub>60</sub> in H<sub>2</sub>O obtained was 8 x 10<sup>-5</sup>M and a 2 γ-cyclodextrin:1 C<sub>60</sub> clathrate is obtained. C<sub>60</sub> is extracted from this aqueous solution with toluene and C<sub>60</sub> of >99 purity is obtained by evaporation. With excess of γ-cyclodextrin more C<sub>60</sub> dissolves and the complex precipitates. The precipitate is insoluble in cold H<sub>2</sub>O but soluble in boiling H<sub>2</sub>O to give a yellow solution. [Andersson et al. *J Chem Soc, Chem Commun* 604 1992.]

(e) C<sub>60</sub> and C<sub>70</sub> can also be readily purified by inclusion complexes with *p*-*tert*-butylcalix[6] and [8]arenes. Fresh carbon-arc soot (7.5g) is stirred with toluene (250ml) for 1 hour and filtered. To the filtrate is added *p*-*tert*-butylcalix[8]arene, refluxed for 10 minutes and filtered. The filtrate is seeded and set aside overnight at 20°. The C<sub>60</sub> complex separates as yellow-brown plates and is recrystallised twice from toluene (1g from 80ml) to give a 90% yield. Addition of CHCl<sub>3</sub> (5ml) to the complex (0.85g) releases C<sub>60</sub> (0.28g, 92% from recrystallised complex).

*p*-*tert*-Butylcalix[6]arene-(C<sub>60</sub>)<sub>2</sub> complex is prepared by adding *p*-*tert*-butylcalix[6]arene (4.4mg) to a refluxing solution of C<sub>60</sub> (5mg) in toluene (5ml). The hot solution is filtered rapidly and cooled overnight to give prisms (5.5mg, 77% yield). Pure C<sub>60</sub> is obtained by decomposing the complex with CHCl<sub>3</sub> as above.

The *p*-*tert*-butylcalix[6]arene-(C<sub>70</sub>)<sub>2</sub> complex is obtained by adding *p*-*tert*-butylcalix[6]arene (5.8mg) to a refluxing solution of C<sub>70</sub> (5mg) in toluene (2ml), filtering hot and slowly cooling to give red-brown needles (2.5mg, 31% yield) of the complex. Pure C<sub>70</sub> is then recovered by decomposing the complex with CHCl<sub>3</sub>. Decomposition of these complexes can also be achieved by boiling a toluene solution over KOH pellets for *ca* 10 minutes. The calixarenes form Na salts which do not complex with the fullerenes. These appear to be the most satisfactory means at present for preparing large quantities of relatively pure fullerene C<sub>60</sub> and C<sub>70</sub> and is considerably cheaper than previous methods. [Atwood et al. *Nature* **368** 229 1994.]

**(f) Rapid and economical purification of C<sub>60</sub> fullerene** [22685-96-8, 99685-96-8] **M 720.6, m >280°, flash point 94°**. Crude soot enriched in fullerenes with ~5% of soluble fullerenes available from suppliers such as *Texas Fullerene Corporation* (8926 Kirby Drive, Houston, USA) and *Polygon Enterprises* (P.O.Box 5536, Waco) was used. A concentrated slurry of fullerene extract (containing *ca* 2.5g of solid) in toluene was prepared by *Soxhlet* extraction of enriched soot with toluene, and evaporated. It was applied onto a glass filter funnel (*D4* 10cm diameter frit) which was loaded with a slurry of silica gel (125g, 230-400 mesh, ~0.05mm, Merck) and acid washed activated charcoal (63g, *Darco G60*, Fluka) in toluene (800ml), and shaken for 5 minutes to give a 5.5 cm high plug which was not allowed to run dry. The plug is covered with 1cm of washed sea sand. The loading of the plug and washing were performed under low vacuum suction (~ 50mm) applied to the suction flask. The plug was rinsed with toluene (3.5L) and the eluate turned intensely purple on further elution with EtOH (300ml). The total separation took ~15 minutes. The purple solution was evaporated to dryness and the residue was washed with small volumes of Et<sub>2</sub>O on a *D4* frit. The yield of C<sub>60</sub> was ~1.5g (60% of the total 75% present in the fullerene soot). Analysis by HPLC should show absence of C<sub>70</sub> or higher fullerenes, but depending upon the original sample may show the presence of 1-3% of C<sub>60</sub>O.

**(g)** A higher purity C<sub>60</sub> can be obtained by using a glass filter funnel with a *D4* frit (diameter 6.5cm) on a suction flask filled with a slurry of silica gel (~84g, 230-400 mesh) and *Darco G60* charcoal (42g) in toluene to a plug height of 9cm; and covered with sea sand (1cm). A concentrated solution of 1.7g of soot extract in toluene (~500ml) was applied onto the plug which was washed with toluene (2.6L). The first fraction (1.4L) of deep purple eluate gave, after evaporation to dryness and washing with small portions of Et<sub>2</sub>O on a *D4* frit, ~690mg of C<sub>60</sub> (41% of the possible 75% of C<sub>60</sub> in the soot extract) which contained 99.75% pure C<sub>60</sub> by HPLC. A further fraction of 1.2L gave, after similar treatment, 143mg (14%) of C<sub>60</sub> of 99.2% purity plus 0.8% of C<sub>60</sub>O as the only other product. [For HPLC and NMR see Isaacs et al. *Helv Chim Acta* **76** 1231 1993].

**(h)** Tour and coworkers [Scrivens et al. *J Am Chem Soc* **114** 7917 1992] described another **gram-quantity purification procedure for C<sub>60</sub> fullerene** where a slurry of alkaline decolorising carbon Norit-A (36g) and silica gel (72g, flash chromatography grade 60 of 230-240 mesh ASTM, 0.040-0.063mm particle size) in toluene (200ml) was poured into a flash chromatography column (38mm in diameter and 45cm long, tape wrapped in case of cracking due to pressure and should be shielded) with a cotton plug at the bottom to form the stationary phase. The slurry was allowed to settle above the stationary phase and drained under a head of 7.5 psi of N<sub>2</sub> pressure applied at the top of the column which should not be allowed to dry out. A saturated solution of crude fullerenes (1.85g, extracted from carbon arc soot) in toluene (400ml, made homogeneous by sonication for ~30 minutes) was applied slowly at the top of the stationary phase. The pressure applied (7.5—10psi) provides a flow rate of ~16ml/minute. C<sub>60</sub> elutes as a purple coloured solution after ~37 minutes in a volume of ~600ml. The eluate, after further ~36 minutes, is colourless and after ~3 minutes further it turns to red-brown when the more strongly adsorbed C<sub>70</sub> fullerene begins to elute out. C<sub>60</sub> fullerene (1.16g, 63 of the 73% of extractable C<sub>60</sub> fullerene) is pure as shown by having one signal *only* in the <sup>13</sup>C NMR (125MHz, \*C<sub>6</sub>H<sub>6</sub>) at δ 143.29 (100%) (compare with the spectrum of C<sub>70</sub> fullerene below).

Repeated chromatography on neutral alumina yields minor quantities of solid samples of C<sub>76</sub>, C<sub>84</sub>, C<sub>90</sub> and C<sub>94</sub> believed to be higher fullerenes. A stable oxide C<sub>70</sub>O has been identified. These have been separated by repeated flash chromatography on alumina with gradient elution using hexane/toluene mixtures (starting from 95:5 and increasing proportions of toluene until the ratio of 50:50 was attained) [Diederich et al. *Science* **252** 548 1991].

**Physical properties of Fullerene C<sub>60</sub>:** C<sub>60</sub> fullerene [135105-52-1] **M 720.64, does not melt below 360°, and starts to sublime at 300° in vacuo**, and is now available commercially in a high state of purity. It is a mustard-coloured solid that appears brown or black with increasing film thickness. It is soluble in common organic solvents, particularly aromatic hydrocarbons which give a beautiful magenta colour. Toluene solutions are

purple in colour. It is soluble in  $^*C_6H_6$  (5mg/ml), but dissolves slowly.  $C_{60}$  crystallises in needles or plates. [Taylor *J Chem Soc, Chem Commun* 1423 1990.]

UV-Vis in hexanes:  $\lambda_{max}$  nm(log  $\epsilon$ ) 211(5.17), 227sh(4.91), 256(5.24), 328(4.71), 357sh(4.08), 368sh(3.91), 376sh(3.75), 390(3.52), 395sh(3.30), 403(3.48), 407(3.78), 492sh(2.72), 540(2.85), 568(2.78), 590(2.86), 598(2.87) and 620(2.60).

IR (KBr):  $\nu_{max}$  1429m, 1182m, 724m, 576m and 527s  $cm^{-1}$ .  $^{13}C$  NMR: one signal with  $\delta$  at 142.68.

$C_{60}$  fullerene is a nano-channel organic semiconductor [Newman et al. *Chem Mater* **16** 4436 2004].

**Purification of [5,6]  $C_{70}$  fullerene for derivatisation** [115383-22-7] **M 840.8, m >280°**. By using a toluene extract (845mg) of Norit A (Aldrich) and a single-column purification using activated charcoal and derivatised polystyrene as stationary phases with toluene/1,2-dichlorobenzene then pure 1,2-dichlorobenzene as eluents, the six fractions obtained (confirmed by HPLC) were: (i) 205mg of pure  $C_{60}$ , (ii) 231mg of 38:62  $C_{60}/C_{70}$ , (iii) 101mg of 88% pure  $C_{70}$ , (iv) 66mg of 92% pure  $C_{70}$ , (v) 131mg of 96% pure  $C_{70}$  and (vi) 57mg of 94% pure  $C_{70}$ . Fractions with >90% purity were satisfactory for the preparation of fullerene  $C_{70}$  derivatives. [Herrmann et al. *Helv Chim Acta* **78** 1673 1995; for column used see Tour above and below.] Fullerene  $C_{70}$  forms a nano adduct with 4,5-dimethoxy-1,2-quinonedimethane [Smith et al. *J Am Chem Soc* **117** 9359 1995, Tour *J Org Chem Perspective* **72** 7477 2007].

Tour and coworkers [Scrivens et al. *J Am Chem Soc* **114** 6939 1994] described a **gram-quantity purification procedure for  $C_{70}$  fullerene** where a slurry of alkaline decolourising carbon Norit-A (400g) and silica gel (800g, flash chromatography grade 60 of 230-240 mesh ASTM, 0.040-0.063mm particle size) in 1:1 toluene/ODCB (*o*-dichlorobenzene) was poured into a flash chromatography column (7cm in diameter and 120cm long, tape wrapped in case of cracking due to pressure and should be shielded) with a cotton plug at the bottom to form the stationary phase. The slurry was allowed to settle above the stationary phase and drained under a head of 15 psi of  $N_2$  pressure applied at the top of the column which should not be allowed to dry out. The stationary phase was 75cm long after settling. A solution of carbon arc soot-extracted [Scrivens & Tour *J Org Chem* **57** 6932 1992] crude fullerene (10.0g, 36%  $C_{70}$  by HPLC analysis) in 1:1 toluene/ODCB (666ml) was applied slowly at the top of the column and eluted with a 15 psi of  $N_2$  head pressure giving a flow rate of ~15ml/minute. Since more solvent was required, the tap at the bottom of the column was closed, the pressure was released, and more solvent was added. After the first colourless fraction (~2516ml), the second purple fraction (1520ml) of  $C_{60}$  fullerene was eluted and gave 5.97g of >99.9% pure  $C_{60}$  fullerene (by HPLC). Then the eluate became colourless again, but was followed by the red-brown colour of  $C_{70}$  fullerene when fraction 3 was collected and the mobile phase was replaced by pure ODCB. Fraction 3 (~750ml) contained a mixture of  $C_{60}/C_{70}$  (58:42, 1.3g) and was followed by a change of colour from red-brown to black indicating that pure ODCB had reached the bottom of the column. The black fraction 4 (750ml) was 90% enriched in  $C_{70}$  fullerene (0.65g), and the final fraction 5 (total 10L) contained 1.58g of  $C_{70}$  of >97% purity. The combined fractions weighed 9.5g (95% mass recovery). Additional column work of impure fractions could give ~3g of >97% pure  $C_{70}$  from 10g of crude material. The solubility of  $C_{70}$  in ODCB in saturated solutions appears to vary with temperature from 24mg/ml at 25° to >50mg/ml in hot ODCB which on standing overnight give an abundance of  $C_{70}$  as long crystals with 99.5% purity (by HPLC). The  $^{13}C$  NMR (125MHz,  $^*C_6H_6$ ) of  $C_{70}$  has  $\delta$  (relative peak heights) at 150.73 (4), 148.18 (11), 147.48 (5), 145.44 (9) and 130.94 (3) (compare with one peak at 142.68 for  $C_{60}$  fullerene previous entry).

**Physical properties of Fullerene  $C_{70}$ :  $C_{70}$  (5,6)-fullerene** [115383-22-7] **M 840.78, does not melt below 360°, and starts to sublime at 300° in vacuo**, and is now available commercially in a high state of purity. It is a reddish-brown solid but greenish black in thicker films. Solutions are port-wine red in colour. Mixtures of  $C_{60}$  and  $C_{70}$  are red due to  $C_{70}$  being more intensely coloured. It is less soluble than  $C_{60}$  in  $^*C_6H_6$  and also dissolves slowly.  $C_{70}$  gives orange-coloured solutions in toluene. Drying at 200-250° is not sufficient to remove all the solvent. Samples need to be sublimed to be free from solvent. [Taylor *J Chem Soc, Chem Commun* 1423 1990.] UV-Vis in hexanes:  $\lambda_{max}$  nm(log  $\epsilon$ ) 214(5.05), 235(5.06), 249sh(4.95), 268sh(4.78), 313(4.23), 330(4.38), 359(4.29), 377(4.45), 468(4.16), 542(3.78), 590sh(3.47), 599sh(3.38), 609(3.32), 623sh(3.09), 635sh(3.13) and 646sh(2.80).

IR (KBr):  $\nu_{max}$  1430 m, 1428m, 1420m, 1413m, 1133mw, 1087w, 795s, 674ms, 642ms, 5778s, 566m, 535ms and 458m  $cm^{-1}$ .

$^{13}C$  NMR [run in the presence of Cr(pentan-2,4-dione)<sub>3</sub> which induces a *ca* 0.12ppm shift in the spectrum]: Five signals with  $\delta$  at 150.07, 147.52, 146.82, 144.77 and 130.28, which are unaffected by proton decoupling.

$C_{70}$  fullerene is a nano-channel organic semiconductor [Newman et al. *Chem Mater* **16** 4436 2004].

**C<sub>76</sub> fullerene** [135113-15-4] **M 912.85, melts > 350°**. It is now available commercially. After the sequential removal of C<sub>60</sub> and C<sub>70</sub> fullerenes from soot extracts (see above) on gel permeation columns (e.g. Buckyclutcher 1 column), C<sub>76</sub> and higher fullerenes are obtained. These are further separated on a Trident-TriDNP functionalised silica column. After two HPLC runs on a C<sub>18</sub> reverse phase (Vydac 201 TP C<sub>18</sub>) column and eluting with 1:1 MeCN/toluene, pure C<sub>76</sub> fullerene is obtained. The identity is confirmed by HPLC/GPC system with Waters 600E UV/VIS detection, mass and NMR spectroscopy. [Seleque et al. In Kadish and Ruoff (Eds) *Fullerenes: Recent Advances in the Chemistry and Physics of Fullerenes and Related Materials* The Electrochemical Soc. Inc, Pennington, NJ, 1994 ISBN 1566770823, Diederich & Whetten *Acc Chem Res* **25** 119 1992, Diederich et al. *Science* **254** 1768 1991.]

**C<sub>78</sub> (C<sub>2v</sub>)-fullerene** [136316-32-0] **M 936.98, melts above 350°**. It is now available commercially. Pure material is obtained as in the previous purification and elutes after C<sub>76</sub> fullerene, followed by **C<sub>78</sub> (D<sub>3h</sub>)-fullerene**. The identities are confirmed by an HPLC/GPC system with Waters 600E UV/VIS detection, mass and NMR spectroscopy. [Seleque et al. In Kadish and Ruoff (Eds) *Fullerenes: Recent Advances in the Chemistry and Physics of Fullerenes and Related Materials* The Electrochemical Soc. Inc, Pennington, NJ, 1994 ISBN 1566770823, Diederich & Whetten *Acc Chem Res* **25** 119 1992, Diederich et al. *Science* **254** 1768 1991, MS and NMR: Taylor et al. *J Chem Soc, Chem Commun* 1043 1992.]

**C<sub>84</sub> fullerene** [135113-16-5] **M 1008.94, melts above 350°**. It is now available commercially. Pure material is obtained as in the previous purification and elutes after C<sub>78</sub> (D<sub>3h</sub>)-fullerene. It consists of at least two isomers. Common impurities are C<sub>82</sub> and C<sub>86</sub> fullerenes. The identities are confirmed by an HPLC/GPC system with Waters 600E UV/VIS detection, mass and NMR spectroscopy. [Seleque et al. In Kadish and Ruoff (Eds) *Fullerenes: Recent Advances in the Chemistry and Physics of Fullerenes and Related Materials* The Electrochemical Soc. Inc, Pennington, NJ, 1994 ISBN 1566770823, Diederich & Whetten *Acc Chem Res* **25** 119 1992, Diederich et al. *Science* **254** 1768 1991, Manolopoulos & Fowler *J Chem Phys* **96** 7603 1992.]

Higher Fullerenes, e.g. **C<sub>396</sub> fullerene** [175833-78-0] have also been isolated [*Chem Abstr* **124** 299339 1996.]

### FUNCTIONALISED FULLERENES.

A large number of functionalised fullerenes have been reported. They are synthesised by chemical reactions at the double bonds of the 'ball' to form anchors for further condensation with small or large molecules such as proteins, lipids, DNA etc. They find applications in a variety of industries including drugs and pharmaceutical industries and in medicine. They are of nano size, can circulate in the animal's body and can be used to deliver substances to targeted tissues. A few of these are commercially available in chemical catalogues and some are described here to show how fullerenes can be made to react with reagents.

**1,4-Bis(pentafluorobenzyl)[C<sub>60</sub>]fullerene {7,8-dihydro-7,8-bis(2,3,4,5,6-pentafluorophenyl)methyl}-[5,6]-fullerene-C60-I<sub>h</sub>}** [1260376-31-5] **M 914.6**. This bifunctional [C<sub>60</sub>]fullerene was prepared in an argon atmosphere by adding potassium metal (124mg, 3.19mmol) in one portion to a freeze-thawed degassed mixture of fullerene (1.0g, 1.93mmol) and 1-methylnaphthalene (5.93g, 4.1.7 mmol, 30 equivalents, see [90-12-0]) in THF (150ml). This produced a dark red solution after stirring at ~25° for 3 hours under argon. Pentafluorobenzyl bromide (3.63g, 13.9mmol, tenfold excess) is then added, stirring is continued for 8 hours, the reaction is stopped by addition of aqueous NH<sub>4</sub>Cl (0.5ml), the mixture is concentrated *in vacuo* down to ~10ml and the crude desired product is precipitated by addition of MeOH. It is purified by chromatography through silica gel and eluting first with CS<sub>2</sub>/hexane (1:1) then CS<sub>2</sub> to provide the *bis-perfluorobenzyl fullerene* (830mg, 55%) which is fully characterised by its spectroscopic properties. It is stable in air and in solution (e.g. CHCl<sub>3</sub>). For further purification preparative HPLC or GPC may be necessary. It is identified by investigations of dynamic light scattering (DLS), X-ray diffraction (XRD), differential scanning calorimetry (DSC), thermogravimetric analysis (TG-DTA), and of its electrochemical (OVP) properties. The data showed that the perfluoroaromatic rings interact in a face-to-face π-π manner with the [C<sub>60</sub>]fullerene surfaces and exhibit unique aggregation-deaggregation behaviour. These properties allow this fullerene to be used for high-performance organic photovoltaic devices. [see Y. Matsuo and coworkers in Li et al. *Chem Commun* **46** 8582 2010, Matsuo et al. *Synfacts* **2** 148 2011.] Note that the name within the chain brackets above is the name given by the *Chemical Abstract Service*.

**tert-Butyl (1,2-methanofullerene C<sub>60</sub>)-61-carboxylate [6,6-closed 1,2-dihydro-1,2-methanofullerene C<sub>60</sub>]-61-carboxylic acid tert-butyl ester** [150493-29-1] **M 834.8**. The crude *tert*-butyl ester was prepared as for the ethyl ester below and obtained in 25% yield as a 1:1:3 mixture of 6,6-closed, trans-5,6-open and cis-5,6-open esters. As for the diethyl ester below, the mixture (20mg) was equilibrated in boiling toluene to give pure 6,6-closed-*tert*-butyl ester (15mg 75%) which held toluene very strongly. It has UV/VIS with  $\lambda_{\max}$  nm( $\epsilon$ ) in toluene at 334 (40250), 406 (*sh* 3850), 429 (1800), 503 (1150), 593 (*br* 750); the IR (KBr) has strong bands at  $\nu_{\max}$  526, 1141, 1733 cm<sup>-1</sup> as well as 5 other medium-weaker bands; the <sup>1</sup>H NMR (500 MHz, in \*C<sub>6</sub>D<sub>6</sub>) has  $\delta$  at 4.38 (*s* 1 H), 1.51 (*s*, 9 H), and the <sup>13</sup>C NMR (125 MHz, CS<sub>2</sub>/(CD<sub>3</sub>)<sub>2</sub>CO 10:1) has  $\delta$  at 164.33(COO), 149.07—71.74 (32 fullerene resonance), 83.28 [COOC(CH<sub>3</sub>)<sub>3</sub>], 41.21 (CHCOO) and 28.58 [(CH<sub>3</sub>)<sub>3</sub>]. The FAB-MS has a peak at 834 (*M*<sup>+</sup>). [Isaacs et al. *Helv Chim Acta* **76** 1231 1993, Isaacs & Dieterich *Helv Chim Acta* **76** 2454 1993.]

**Diethyl (1,2-methanofullerene C<sub>60</sub>)-61,61-carboxylic acid [6,6-closed-1,2-dihydro-1,2-methanofullerene C<sub>60</sub>]-61,61-carboxylic acid diethyl ester** [155679-98-4] **M 878.8, m >270°**. If the <sup>1</sup>H NMR spectrum indicated the presence of isomeric esters then dissolve it in chlorobenzene (0.2g in 100ml) and heat it under reflux for 24 hours to convert the isomers to the more stable [6,6-closed]-ester. Column chromatography on silica and elution first with toluene/hexane (1:1) then toluene followed by recrystallisation from CHCl<sub>3</sub>/MeOH and drying (60°/0.1 Torr, 6 hours) provided the desired diester as a dark solid [*R*<sub>F</sub>: SiO<sub>2</sub> toluene is 0.50]. Store it at 2-8°. It has IR (KBr) with  $\nu_{\max}$  bands at 1061w, 1095m, 1186 (C<sub>60</sub>), 1206m, 1234s, 1266m, 1295m, 1428 (C<sub>60</sub>), 1745 (C=O) and 2979w cm<sup>-1</sup>. The <sup>1</sup>H NMR (200MHz; CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 4.75 (q, *J* = 7Hz) and 1.53 (t, 6H, *J* = 7Hz); and the <sup>13</sup>C NMR (125MHz; CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 163.61, 145.33—139.04 (16 peaks), 71.61, 63.42, 52.21, 14.22. The *di-tert-butyl ester* was similarly purified (see above). [Bingel *Chem Ber* **126** 1957 1993, Dieterich et al. *J Chem Soc, Perkin Trans II* 391 1994.]

**Diethyl (1,2-methanofullerene C<sub>70</sub>)-71,71-carboxylate [6,6-closed-1,2-dihydro-1,2-methanofullerene C<sub>70</sub>]-71,71-carboxylic acid diethyl ester** [153218-95-2] **M 998.9**. It was prepared by reaction of diethyl bromomalonate with fullerene C<sub>70</sub> in the presence of DBU in toluene (4 hours, ~25°), filtered and purified by chromatography through Kieselgel (0.063—0.2mm) by eluting with a gradient of toluene/*iso*-hexane (1/1 to 4/1). This gave the *diester* in 60% yield [*R*<sub>F</sub>: SiO<sub>2</sub> toluene is 0.50], with a 40% recovery of the fullerene. Store it at 2-8°. The <sup>1</sup>H NMR (360MHz; CDCl<sub>3</sub>) has  $\delta_{\text{H}}$  at 4.75 (q, *J* = 7.1Hz) and 1.46 (t, 6H, *J* = 7.1Hz); and the <sup>13</sup>C NMR (100MHz; CDCl<sub>3</sub>) has  $\delta_{\text{C}}$  at 163.45, 155.33—130.83 (34 peaks), 66.90, 66.24, 63.47, 37.22, 14.23. [Bingel *Chem Ber* **126** 1957 1993.]

**1,2-Dihydro-[1,2]fullereneC<sub>60</sub>[3,4]pyrrolidin-1-ylethoxyethoxyethylammonium bis-trifluoroacetate {2-[2-[2-(2'*H*-[5,6]fullereno-C<sub>60</sub>-I<sub>h</sub>-[1,9-*c*]pyrrol-1'(5'*H*)-yl)ethoxy]ethoxy]ethanamine bis-2,2,2-trifluoroacetate}** [C<sub>60</sub>=CH)<sub>2</sub>NH<sup>+</sup>-CH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>NH<sub>3</sub><sup>+</sup>. 2CF<sub>3</sub>COO<sup>-</sup>, C<sub>70</sub>H<sub>20</sub>O<sub>6</sub>N<sub>2</sub>F<sub>6</sub>] [356066-52-9 for mono(trifluoroacetate)] **M 1123.0**. *N*-Boc-aminoethoxyethoxyethylaminoacetic acid (423mg, 1.38mmol, see below in SWCNT[=CH)<sub>2</sub>N-CH<sub>2</sub>CH<sub>2</sub>O-CH<sub>2</sub>CH<sub>2</sub>O-CH<sub>2</sub>CH<sub>2</sub>NH<sub>3</sub><sup>+</sup>. HCl]<sub>n</sub>), paraldehyde (456mg, 3.45mmol) and fullerene C<sub>60</sub> (500mg, 0.69mmol) in toluene (300ml) are refluxed for 1 hour. After cooling to ~25° the mixture is applied onto a column of silica gel (NM Kieselgel 60, 70-320 mesh) and eluted with 9:1 toluene/EtOAc to give the analytically pure (elemental C, H, N) C<sub>60</sub>=*N*-Boc-pyrrolidinyl derivative (212mg, 0.29mmol, 31%). It has IR-DRIFT (KBr) with bands at  $\nu_{\max}$  545, 794, 1121, 1168, 1457, 1498, 1710, 2853, 2921, 3444 cm<sup>-1</sup>; the UV-VIS (cyclohexane) has  $\lambda_{\max}$  at 254, 309, 429, 466sh, 546sh, 703 nm; the <sup>1</sup>H NMR (200 MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 5.08 (bs, 1H), 4.50 (s, 4H), 4.05 (t, *J* = 5.5Hz, 2H), 3.77 (m, 4H), 3.60 (t, *J* = 5.3Hz, 2H), 3.36 (m, 4H), 1.44 (s, 9H); the <sup>13</sup>C NMR (50 MHz, CDCl<sub>3</sub>, TMS) has  $\delta$  at 156.0, 155.1, 147.3, 146.3, 146.1, 145.7, 145.4, 145.3, 144.6, 143.1, 142.7, 142.3, 142.1, 141.9, 140.2, 136.2, 79.3, 70.9, 70.7, 70.6, 70.5, 68.6, 54.4, 28.6; and the ES-MS showed a peak at *m/z* 995 (MH<sup>+</sup>).

The *N*-Boc group is removed by treating the pure C<sub>60</sub>=*N*-Boc-pyrrolidinyl derivative (150mg, 0.14mmol) in CH<sub>2</sub>Cl<sub>2</sub> (3ml) with CF<sub>3</sub>COOH (3ml), and stirring for 3 hours. The solvent is removed *in vacuo*, the residue is washed with toluene and dried *in vacuo* to give the analytically pure (elemental C, H, N) functionalised title fullerene C<sub>60</sub>=*N*-pyrrolidinyl bis-trifluoroacetate salt (168mg, 0.15mmol, ~99%). This has IR-DRIFT (KBr) with bands at  $\nu_{\max}$  525, 715, 1130, 1190, 1480, 1515, 1680, 2890 cm<sup>-1</sup>; UV-VIS (cyclohexane) with  $\lambda_{\max}$  at 331, 430, 484sh, 684 254, 309, 429, 466sh, 546sh, 703 nm; the <sup>1</sup>H NMR (200 MHz, DMSO-*d*<sub>6</sub>, TMS) has  $\delta$  at

7.81 (brs, 3H), 4.58 (s, 4H), 3.93 (t,  $J = 5.5\text{Hz}$ , 2H), 3.71-3.43 (m, 6H), 3.32 (m, 2H), 2.95 (m, 2H); the  $^{13}\text{C}$  NMR (50 MHz, DMSO- $d_6$ , TMS) has  $\delta$  at 156.0, 147.4, 146.7, 146.4, 146.2, 146.0, 145.5, 145.4, 144.7, 143.3, 142.8, 142.5, 142.2, 142.0, 140.2, 136.4, 135.4, 71.5, 70.5, 70.1, 68.0, 67.5, 54.1; and the ES-MS showed a peak at  $m/z$  896 ( $\text{MH}^+$ ). Related compounds with a two carbon atoms and a thirteen carbon atom chain between the pyrrolidine nitrogen atom and the terminal amino nitrogen atom have also been prepared. The  $\omega$ -primary amino groups at the end of the chains were derivatised (e.g. with 12-acetylsulfanildodecanoic acid in the presence of EDCI and HOBt for self-assembled monolayers (SAMs) purposes, or various fluorescent indole-2-carboxylic acids) by standard procedures. Useful applications for nano materials science and nano medicinal chemistry (e.g. by linking to DNAs or other biological materials) have been made possible with such fullerene precursors by M. Prato and coworkers. [Kordatos et al. *J Org Chem* **66** 4915 2001.]

**Ethyl (1,2-methanofullerene  $\text{C}_{60}$ )-61-carboxylate [6,6-closed 1,2-dihydro-1,2-methanofullerene  $\text{C}_{60}$ ]-61-carboxylic acid ethyl ester** [50493-27-9] **M 806.8**. Fullerene  $\text{C}_{60}$  reacts with an equimolar amount of ethyl diazoacetate in boiling toluene for 7 hours to provide a 30% yield of a 1:1:3 mixture of 6,6-closed : trans-5,6-open : cis-5,6-open *tert*-butyl (1,2-methanofullerene  $\text{C}_{60}$ )-61-carboxylate. This mixture was isolated by applying the residual solid from the reaction onto a silica gel column which was first eluted with hexane to remove fullerene  $\text{C}_{60}$  then with toluene-hexane (1:1) to provide the mixture at  $R_F$  ca 0.43. The mixture was isomerised almost entirely to the more stable 6,6-close ester by refluxing it (35mg) in toluene (105ml) for 24 hours when the initial purple solution turned pink in colour. Chromatography on silica gel as before provided a pink-red product eluting at the same  $R_F$  ca 0.43 which was then evaporated to dryness. The black microcrystalline residue was washed with  $\text{Et}_2\text{O}$  and dried at  $25^\circ/0.1$  Torr (31mg, 89%). It holds residual toluene tenaciously. It has UV/VIS with  $\lambda_{\text{max}}$  nm( $\epsilon$ ) in toluene at 331 (40250), 395 (*sh* 5250), 404 (*sh* 3300), 417 (1900), 429 (2250), 495 (*br* 1250); the IR (KBr) has strong bands at  $\nu_{\text{max}}$  527, 1155, 1185, 1742  $\text{cm}^{-1}$  as well as 6 other medium to weak bands; the  $^1\text{H}$  NMR (400 MHz, in  $^*\text{C}_6\text{D}_6$ ) has  $\delta$  at 4.31 (*s* 1 H), 4.10 (*q*,  $J = 7.1\text{Hz}$ , 2 H), 1.06 (*t*,  $J = 7.1\text{Hz}$ , 3 H), and the  $^{13}\text{C}$  NMR (125 MHz,  $\text{CS}_2/(\text{CD}_3)_2\text{CO}$  10:1) has  $\delta$  at 165.44 ( $\text{COOEt}$ ), 149.07—71.59 (32 fullerene resonance), 62.70 ( $\text{CH}_3\text{CH}_2$ ), 40.01 ( $\text{CHCOOEt}$ ) and 14.99 ( $\text{CH}_3\text{CH}_2$ ). The FAB-MS has an  $m/z$  peak at 807 ( $M^+$ ), and the elemental analysis is consistent with  $\text{C}_{62}\text{H}_6\text{O}_2$ . [Isaacs et al. *Helv Chim Acta* **76** 1231 1993, Isaacs & Dieterich *Helv Chim Acta* **76** 2454 1993.]

**Fullerene  $\text{C}_{60}$ - $\gamma$ -lactone esters.** G-W. Wang and coworkers [Li et al. *Org Lett* **21** 4896 2010] have improved the condensation yields of diethyl malonate esters with [ $\text{C}_{60}$ ]fullerene to form the disubstituted fused  $\gamma$ -lactone esters by using  $\text{Fe}(\text{ClO}_4)_3$  instead of  $\text{Mn}(\text{OAc})_3$  for assistance (catalysis). Although the yields are still low, they are however, considerably higher when mediated by the iron perchlorate, and the products are cleaner. The solvent is *o*-dichlorobenzene and the ratios of reagents are 1:2:2:20 for [ $\text{C}_{60}$ ]fullerene:EtOCOCHR $\text{COOEt}$ : $\text{Fe}(\text{ClO}_4)_3$ : $\text{Ac}_2\text{O}$ . The following results were obtained: **R** ester (reaction temperature, reaction time, **yield** of  $\gamma$ -lactone ester, consumed fullerene): **Me** ( $80^\circ$ , 20 minutes, **34%**, 55%); **Et** ( $80^\circ$ , 20 minutes, **37%**, 71%); **PhCH<sub>2</sub>** ( $80^\circ$ , 30 minutes, **27%**, 61%); **Ph\*** ( $0^\circ$ , 180 minutes, **12%**, 67%); **Br** ( $80^\circ$ , 30 minutes, **22%**, 56%); and **EtOCO** ( $110^\circ$ , 20 minutes, **16%**, 53%); \* reactant ratio was 1:2:2:50. The structures of the lactones are consistent with their HRMS,  $^1\text{H}$  NMRs,  $^{13}\text{C}$  NMRs, FT-IRs and UVs. It should be possible to convert these lactones into hydrofullerenes, fullerene hemiacetals, fullerene hemiketals and fullerenols. [Li et al. *Synfacts* **1** 38 2011.]

**Fullerene  $\text{C}_{60}$ /Poly(bisphenol A)carbonate.**  $\text{C}_{60}$  is a known 'radical sponge' [Morton et al *J Am Chem Soc* **114** 5454 1992, McEwen et al. *J Am Chem Soc* **114** 4412 1992]. It reacts readily with free radicals and undergoes photolysis to generate  $\text{C}_{60}$  radical ions. Thus irradiating a mixture of  $\text{C}_{60}$  and poly(bisphenol A)carbonate polymer (PC, [25037-45-0]) at room temperature with a conventional UV lamp, or warming the mixture at  $60^\circ$  with AIBN (a radical initiator, see [78-61-1]) results in fullerenated-PC. By controlling the ratio of reagents (i.e. 5.4mg  $\text{C}_{60}$ /500mg PC in 1,1,2,2-tetrachloroethane), a  $\text{C}_{60}$ /PC can be obtained with a  $\text{C}_{60}$  content as high as 6.3% (i.e.  $>2$   $\text{C}_{60}$  per PC chain) in  $\sim 99\%$  yield. Multi-additions of PC did not lead to heavy cross-linking. The fullerenated polymer is a brown powder which is soluble in common organic solvents such as THF and  $\text{CHCl}_3$ . When the polymer is dissolved in THF (a solvent that does not dissolve  $\text{C}_{60}$ ) and the solution is precipitated into hexane through a filter, no particles are left on the filter, and the hexane supernatant is colourless. This shows that all the  $\text{C}_{60}$  is incorporated into the polymer. The precipitate is then collected and dried under high vacuum to constant weight. The  $\text{C}_{60}$  content is 1.16 wt%, has  $M_n \sim 27,000$ ; the UV (0.3mg/ml in THF) has  $\lambda_{\text{max}}$  at 238.9, 259.4, 262.0, 287.2, 329.5sh nm, and the IR (KBr) has  $\nu_{\text{max}}$  at 3040w, 2968s,

1778vs, 1602m, 1506s, 1410m, 1386w, 1364m, 1230vs, 1194s, 1162s, 1080s, 1014s, 888m, 830s, 768m, 556s, 528w  $\text{cm}^{-1}$ . This fullerenation provides a versatile synthetic tool for making processable fullerene polymers and opens opportunities for exploring fullerene-based speciality materials with properties not found in the unconjugated polymers. [Tang et al. *Macromolecules* **30** 2848 1997.]

**(1,2-Methanofullerene  $\text{C}_{60}$ )-61-carboxylic acid [6,6-closed-1,2-dihydro-1,2-methanofullerene  $\text{C}_{60}$ ]-61-carboxylic acid** [155116-19-1] **M 778.7**. Attempted hydrolysis of the ethyl ester (see [155679-98-4] above) was unsuccessful, however, treatment of the 6,6-closed-ethoxycarbonylmethyl ester with  $\text{BBr}_3$  in  $^*\text{C}_6\text{H}_6$  under  $\text{N}_2$  and stirring for 9 hours gave an 82% yield of the acid. Similarly the *tert*-butyl ester (see [150493-29-1] above), but in refluxing toluene for 8 hours, gave a 77% yield of the acid. After hydrolysis, the acid was purified by dissolving it in  $\text{CHCl}_3/\text{Me}_2\text{SO}$ , re-precipitating with hexane and drying overnight at  $25^\circ/0.1$  Torr, or at  $60^\circ/0.1$  Torr for 12 hours. The brown or black solid 6,6-closed-acid was mostly insoluble in common solvents and slightly soluble in bromobenzene and 1,2-dichlorobenzene. This 6,6-closed acid (i.e. cyclopropane acid) has IR (KBr) with strong bands  $\nu_{\text{max}}$  at 525, 699, 810, 847, 1013, 1426, 1785, 1794  $\text{cm}^{-1}$  and twenty other medium to weak bands. The  $^1\text{H}$  NMR (300 MHz, in  $\text{CHCl}_3/\text{Me}_2\text{SO}$ , 1:1) has one signal for the cyclopropane H at 5.13 ppm, the  $^{13}\text{C}$  NMR consists of 32 signals between  $\delta$  136 and 167, and signals at  $\delta$  40.34 and 71.66. Elemental analysis is consistent with the formula  $\text{C}_{62}\text{H}_2\text{O}_2 \cdot 0.5\text{Me}_2\text{SO}$ .

With alcohols in the presence of dicyclohexylcarbodiimide, 1-*H*-benzotriazol-1-ol and a base [e.g. 4-( $\text{Me}_2\text{N}$ ) $\text{C}_6\text{H}_4\text{N}$  or  $\text{Et}_3\text{N}$ ] in bromobenzene, the corresponding esters were obtained. Similarly peptides with *methyl glycinate* and *methyl-L-phenylalaninate* were obtained by using the corresponding amino acid esters. [Isaacs & Diederich *Helv Chim Acta* **76** 2454 1993, Diederich et al. *Chem Soc Rev* **23** 243 1994.]

**6,9,12,15,18-Pentamethyl-1,6,9,12,15,18-hexahydro( $\text{C}_{60}$ -I<sub>h</sub>)fullerene** [244229-54-7] **M 796.7**. Preparation of these substituted fullerenes should be carried out in Schlenk-type equipment under an inert atmosphere ( $\text{N}_2$  or argon), in dry solvents and reagents (at least in the early stages of the reactions), and any oxygen or air in solvents should be flushed out by bubbling dry inert gas through them. Microcrystalline [ $\text{C}_{60}$ ]fullerene (2.0g, 2.78mmol, [99685-96-8]) in 1,2-dichlorobenzene (90ml), cooled in an ice-water bath, stirred under reduced pressure (1mm) to remove dissolved  $\text{O}_2$ , and warmed to  $\sim 25^\circ$ , is added during 15 minutes to a stirred mixture of  $\text{CuBr} \cdot \text{SMe}_2$  (6.84g, 33.3mmol, use an efficient fume cupboard) in THF (47ml) at  $\sim 25^\circ$  to which is previously added  $\text{MeMgBr}$  in THF (3 M, 11.1ml, 33.3mmol) followed by DMI (3.62ml, 33.3mmol, see [80-73-9]) and rapidly warmed to  $\sim 35^\circ$  within 5 minutes. The colour of the white suspension soon turned to dark brown. After stirring at  $35^\circ$  for 40 minutes under a flow of  $\text{N}_2$  (note that some liberation of ethane may occur), a degassed saturated aqueous  $\text{NH}_4\text{Cl}$  solution (3.0ml) is added *via* a syringe, the colour of the solution changes from dark brown to red-brown and the mixture is stirred under a vacuum (*ca* 1mm) at  $\sim 25^\circ$  to remove THF and  $\text{Me}_2\text{S}$  and reduce it to half its volume. This is diluted with degassed toluene (200ml) and subjected to silica gel flash chromatography (45 x 200mm size, using 90g of silica gel 230-450 mesh) with toluene as eluent (total volume 100ml). It should be done as rapidly as possible, as care should be taken to avoid oxygen and formation of  $\text{C}_{60}\text{Me}_5\text{O}_n\text{H}$  ( $n = 2-3$ ). The vermilion eluate is evaporated at  $\sim 40^\circ/10\text{mm}$  then at  $80^\circ$  (to remove 1,2-dichlorobenzene) until solid begins to appear on the sides of the flask.  $\text{N}_2$  is allowed to enter the evacuated flask and degassed MeOH is added along the inside wall of the flask, whereby the mixture becomes cloudy, and MeOH ( $\sim 400\text{ml}$  total) is added until precipitation is complete. The solid is filtered off, washed with MeOH (5 x 10ml) and hexane (3 x 10ml) and dried *in vacuo* (1mm) to give red microcrystals of pentamethylfullerene (2.08g, 94%) of  $\sim 91\%$  purity. Purity and/or purification is carried out using a Buckyprep or ODS column (4.6 x 150mm, flow rate 1ml/minute) and eluting with toluene/*i*-PrOH (7:3) or (3:7) respectively. The solid should be stored in an inert atmosphere, as on storage in air it deteriorates slowly over several months, and in solution in the presence of air its purity decreases by 80% in 24 hours. It has IR (neat) with  $\nu_{\text{max}}$  bands at 2973, 2912, 2858, 1572, 1546, 1518, 1459, 1370, 1324, 1286, 1265, 1234, 1169, 1145, 1128, 792, 747, 684  $\text{cm}^{-1}$ ; the  $^1\text{H}$  NMR (500 MHz,  $\text{CS}_2/(\text{CD}_3)_2\text{CO}$  5%) has  $\delta$  at 2.34 (s, 6H), 2.35 (s, 6H), 2.46 (s, 3H), 4.50 (s, 1H,  $\text{C}_{60}\text{-H}$ ); and the  $^{13}\text{C}$  NMR (100 MHz,  $\text{CS}_2/(\text{CD}_3)_2\text{CO}$  5%) has  $\delta$  at 27.23 (2C), 27.61 (2C), 33.14 (1C), 51.37 (1C), 51.45 (2C), 53.51 (2C), 59.75 (1C), 143.61 (2C), 144.12 (2C), 144.48 (2C), 144.53 (2C), 144.77 (2C), 144.89 (4C), 145.20 (2C), 145.42 (2C), 145.87 (2C), 145.96 (2C), 146.98 (1C), 147.14 (2C), 147.24 (2C), 148.03 (2C), 148.12 (2C), 148.28 (3C), 148.54 (2C), 148.60 (2C), 148.91 (4C), 149.10 (2C), 150.12 (2C), 154.27 (2C), 154.37 (2C), 154.62 (2C), 157.76 (2C). [Matsuo, Mueamatsu, Tahara, Koide and Nakamura *Org Synth* **83** 80 2006, Matsuo et al. *Functional Organic Materials* 58-80 2007.]

The procedure described above is unique as all five substituents are symmetrically placed on the bridgehead carbons of five six-membered rings surrounded by a fused five membered cyclopentadiene ring (X-ray evidence). Similar reactions also produce very high yields of fullerenes with *five*  $C_6H_5$ , *five*  $4-CF_3C_6H_5$ , *five*  $4-MeOC_6H_4$ , *five*  $4-ClC_6H_5$ , *five*  $4-PhC_6H_5$ , *five* (*E*)-1-propenyl, *five* (*Z*)-1-propenyl and *five* (*E*)-2-phenylethenyl groups all symmetrically placed as in the pentamethyl derivative above. [see Matsuo, Mueamatsu, Tahara, Koide and Nakamura *Org Synth* **83** 80 2006 and references therein.]

The hydrogen atom of the central cyclopentadiene ring can be displaced by potassium with *t*-BuOK to form the  $K(C_{60}Me_5)$  complex in which the potassium atom can then be displaced to form iridium complexes such as  $Ir(\eta^5-C_{60}Me_5)(CO)_2$  [Matsuo, Iwashita and Nakamura *Organometallics* **24** 89 2005], and with rhodium to form  $Rh(\eta^5-C_{60}Me_5)(CO)_2$  [Sawamura, Kuninobu and Nakamura *J Am Chem Soc* **122** 12407 2000], structures which are supported by X-ray structure analyses. The same cyclopentadiene can complex with Fe and cyclopentadiene (Cp) to form a hybrid of 'buckyferrocene'  $Fe(C_{60}Me_5)Cp$  [Sawamura et al. *J Am Chem Soc* **124** 9354 2002, Nakamura *Pure Appl Chem* **75** 427 2003].

***N*-Tosyl[1,2]-aziridino[ $C_{60}$ ]fullerene {2a[(4-methylphenyl)sulfonyl]-2a-aza-1,2(2a)-homo[5,6]fullerene- $C_{60}$ -I<sub>h</sub> [226909-63-3] M 889.0.** This useful fullerene precursor is readily prepared. In a dry Schlenk flask with flushing argon is added CuCl (1.2mg, 12μmol), *o*-dichlorobenzene (40ml) and 2,6-toluidine (2.0 μL, 24 μmol), then stirring at ~25° for 30 minutes, followed by adding  $C_{60}$  fullerene (432mg, 600μmol) and TsN=IPh (224mg, 600μmol, see synthesis below\*). The mixture is stirred ~25° for 12 hours and purified by flash chromatography through silica gel (toluene/hexane 1.2 ~ 1.1) to provide the pure aziridinofullerene (228mg, 43%) as a dark brown solid [and recovered 185mg, 43%, of unreacted  $C_{60}$  fullerene]. It has <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>/CS<sub>2</sub> = 1:1, TMS) with δ at 2.56 (s, 3H), 7.50 (d, *J* = 8.3Hz, 2H), 8.19 (d, *J* = 8.3Hz, 2H); the <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>, CHCl<sub>3</sub> at δ 77.0 standard) has δ at 21.75, 79.69, 128.44, 129.93, 135.63, 140.75, 141.20, 141.71, 142.02, 142.63, 142.94, 142.99, 143.12, 143.74, 143.81, 144.01, 144.36, 144.83, 144.89, 144.97, 145.14, 145.34; and the HRMS (ESI-TOF; negative) showed a peak at *m/z* 889.0205, and 889.0203 is the calculated value for  $C_{67}H_7NO_2S$  (M<sup>-</sup>). Note that other Cu catalysts and other pyridine bases resulted in lower yields of this reaction. Itami and coworkers [Nambo et al. *J Am Chem Soc* **133** 2402 2011] developed the use of this aziridinofullerene as a versatile platform for preparing a variety of functionalised fullerenes. Thus it reacts with aryl and heteroaryl compounds (5% TfOH, 1,2-Cl<sub>2</sub>C<sub>6</sub>H<sub>4</sub>, 100°, 12 hours) to replace the 1,2-aziridine by two aryl or heteroaryl groups to form 1,4-disubstituted fullerenes in over 80% yields; it reacts with bifunctional nucleophiles of the type R-Z(H)=Y(X-H)-R' (10% TfOH, 1,2-Cl<sub>2</sub>C<sub>6</sub>H<sub>4</sub>, 100°, 12 hours) to form 1,2-disubstituted fullerenes (1,2-disubstituted fullerenes fused with a 5-membered ring where the H's are replaced by bonds with C1 and C2 of the fullerene), such as 2,3-furano, 2,3-pyrrolo, 2,3-thiazolo-, 1,2-cyclic [1,3,2]-dioxaborolan- fullerenes in over 60% yields; it reacts with 2,2'-bithiophenes in various ways, and undergoes formal [2+2] cycloaddition reactions with 1,2-[bis(*p*-methoxyphenyl)acetylene (10% TfOH, 1,2-Cl<sub>2</sub>C<sub>6</sub>H<sub>4</sub>, 100°, 12 hours) to form a 3,4-adduct (cyclobutene) which retains the 1,2-aziridino moiety in 81% yields. Some of the structures were supported by X-ray crystal structure analysis.

\*[*N*-(*p*-Toluenesulfonyl)imino]phenyliodinane (TsN=IPh) [55962-05-5] M 373.1. Toluene-*p*-sulfonamide (3.42g, 20mmol) and diacetoxyiodobenzene (6.4g, 20mmol, see [2340-34-4]) are added to a solution of KOH (2.8g, 50mmol) dissolved in MeOH (80ml) at 0°, and then stirred at 25° for 3 hours. The mixture is poured into distilled water at 0°, kept at 4° overnight, the solid is filtered off, drained and recrystallised from hot MeOH (20ml) to give TsN=IPh (2.6g, 35%). It has <sup>1</sup>H NMR (200 MHz, DMSO-*d*<sub>6</sub>, TMS) with δ at 7.80-7.73 (m, 2H), 7.56-7.48 (m, 3H), 7.40-7.31 (m, 2H), 7.15-7.10 (m, 2H), 2.34 (s, 3H). [Heuss et al. *Inorg Chim Acta* **342** 301 2003.]

**2,5,10-Triphenyl-2,5,6,10-tetrahydro( $C_{70}$ -I<sub>h</sub>)fullerene [ $C_{70}$ (4-PhC<sub>6</sub>H<sub>4</sub>)<sub>3</sub>H] [244237-40-9] M 1073.0.** When a procedure similar to the preparation of  $C_{60}Me_5H$  above is applied to fullerene- $C_{70}$  only three substituents are inserted. A suspension of CuBr.SMe<sub>2</sub> (374mg, 1.82mmol, 30 equivalents) in THF (23.0ml) is treated with PhMgBr (0.98M, 1.86ml, 1.82mmol, 30 equivalents) and stirred at 28° for 20 minutes. To the resulting yellow suspension is added a degassed solution of  $C_{70}$  (49.6mg, 59μmol) in 1,2-dichlorobenzene (25ml) and stirring is continued for 24 hours. The reaction is treated with 5% aqueous HCl, the organic layer is washed with H<sub>2</sub>O, brine, dried (MgSO<sub>4</sub>), evaporated to a small volume and diluted with MeOH to give a dark brown precipitate which is washed thoroughly with MeOH, Et<sub>2</sub>O, H<sub>2</sub>O again and dried *in vacuo* to give  $C_{70}Ph_3H$  (61.2mg, 95%,

96% purity, cf C<sub>60</sub>Me<sub>5</sub>H above). The <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>) has δ at 7.8-7.78 (m, 2H), 7.74-7.70 (m, 2H), 7.61-7.57 (m, 3H), 7.39-7.23 (m, 8H), 4.43 (s, 1H); and the <sup>13</sup>C NMR (100 MHz, CDCl<sub>3</sub>-CS<sub>2</sub>) has δ at 160.70, 155.67, 155.07, 153.80, 152.62, 153.14, 152.62, 152.19, 150.16, 149.87, 149.57, 149.42, 149.39, 149.17, 149.08, 148.95, 148.82, 148.68, 148.59, 148.30, 148.21, 148.09, 147.92, 147.88, 147.77, 147.68, 147.18, 147.06, 147.00, 146.65, 146.61, 146.44, 146.23, 146.18, 145.84, 145.64, 145.51, 145.32, 145.06, 145.03, 144.92, 144.88, 144.74, 144.48, 144.15, 143.97, 143.43, 142.99, 142.81, 142.21, 142.18, 140.50, 140.40, 139.54, 139.20, 138.33, 136.87, 133.08, 131.96, 131.93, 131.82, 131.72, 131.59, 131.43, 131.43, 131.41, 130.65, 128.96, 128.88, 128.78, 128.16, 128.04, 127.71, 127.60, 127.47, 127.27, 127.15, 126.70, 126.42, 56.23, 56.19, 55.5; and the APCI-MS has *m/z* 1072 [(M - H)<sup>-</sup>]. Similarly prepared are C<sub>70</sub>Me<sub>3</sub>H, C<sub>70</sub>(4-PhC<sub>6</sub>H<sub>4</sub>)<sub>3</sub>H, C<sub>70</sub>(4-ClC<sub>6</sub>H<sub>4</sub>)<sub>3</sub>H, and C<sub>70</sub>(1-naphthyl)<sub>3</sub>H. [Sawamura et al. *J Mater Chem* **12** 2109 2002, cf Matsuo, Mueamatsu, Tahara, Koide and Nakamura *Org Synth* **83** 80 2006 and references therein.] C<sub>70</sub>(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>3</sub>H is also prepared similarly and is believed to proceed by formation of [(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>2</sub>Cu<sup>-</sup>] from the reaction between the arylMgBr and CuBr.SMe<sub>2</sub> which attacks successively three six-membered rings around a cyclopentadiene ring to give the intermediate C<sub>70</sub>(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>3</sub>Cu-Ar which decomposes on addition of acid or NH<sub>4</sub>Cl to provide the desired C<sub>70</sub>(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>3</sub>H. When the latter, in THF-*d*<sub>8</sub>, is treated with 1 equivalent of *t*-BuOK or TIOEt at 25°, the colour changes from reddish-brown to dark red to provide K[η<sup>5</sup>-C<sub>70</sub>(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>3</sub>] {whose <sup>1</sup>H NMR (400 MHz, THF-*d*<sub>8</sub>) has δ at 8.07-8.05 (m, Ar-H, 6H), 7.57-7.54 (m, ArH, 6H); and the <sup>13</sup>C NMR (100 MHz, THF-*d*<sub>6</sub>) has δ at 166.18, 160.50, 153.30, 151.99, 151.04, 150.32, 149.79, 149.77, 149.17, 148.88, 148.82, 148.48, 148.44, 148.26, 147.14, 146.90, 146.81, 146.37, 146.19, 146.07, 145.72, 145.61, 145.43, 144.71, 142.53, 142.28, 137.87, 136.81, 143.54, 134.46, 133.24, 132.92, 132.60, 129.36-129.23 (m), 128.93, 125.80-125.77 (m), 125.6 (q, *J* = 271.19Hz), 121.07, 61.06, 59.11}, or Tl[η<sup>5</sup>-C<sub>70</sub>(4-CF<sub>3</sub>C<sub>6</sub>H<sub>4</sub>)<sub>3</sub>] {whose <sup>1</sup>H NMR (400 MHz, THF-*d*<sub>8</sub>) has δ at 8.09-8.07 (m, Ar-H, 6H), 7.59-7.53 (m, ArH, 6H)} respectively [Sawamura, Iikura, Hirai and Nakamura *J Am Chem Soc* **120** 8285 1998].

**1,4,11,15,30-Pentakis(4-hydroxyphenyldimethylsilylmethyl)-2H-1,2,4,11,15,30-hexahydro-(60)fullerene {4-HOC<sub>6</sub>H<sub>4</sub>Si(Me)<sub>2</sub>CH<sub>2</sub>]<sub>5</sub>-2H-1,2,4,11,15,30-hexahydro-(C<sub>60</sub>-I<sub>h</sub>)fullerene, 1,7,8,11,24,27-hexahydro-1,7-, 11,24,27-pentakis(dimethyl 4-hydroxyphenylsilyl)methyl-[5,6]fullerene-C<sub>60</sub>-I<sub>h</sub>}** [658080-04-7] **M 1549.1.**

This functionalised fullerene can be used as an example of a molecule to which long chains can be coupled for specialised purposes. Typically, using Schlenk equipment under N<sub>2</sub> or argon in strict absence of H<sub>2</sub>O, the protected Grignard compound Me<sub>2</sub>(4-tetrahydropyranyloC<sub>6</sub>H<sub>4</sub>)SiCH<sub>2</sub>MgCl (0.656M in THF, 13.0ml, 9.84mmol) is added to a purple suspension of a white powder of CuBr.SMe<sub>2</sub> (3.00g, 14.6mmol) and microcrystalline [C<sub>60</sub>]fullerene (400mg, 0.555mmol) in 1,2-dichlorobenzene (80ml), and stirred at 25° for 1 hour. The mixture is treated with saturated aqueous NH<sub>4</sub>Cl (0.5ml), concentrated *in vacuo* to ca 60ml and diluted with toluene. The brown mixture is applied onto a silica column and washed with toluene (~100ml, i.e. until the effluent becomes yellow), followed by EtOAc/toluene (3:97) and the effluent is collected until its colour is pale yellow. This is evaporated *in vacuo* to a small volume and MeOH is added rapidly to precipitate a brown powder which is collected under N<sub>2</sub>, dissolved in MeOH/toluene (1:1, 46ml) and TsOH.H<sub>2</sub>O (92mg, 0.48mmol) is added to it. After stirring for 1 hour at 25°, the mixture is neutralised with NaHCO<sub>3</sub>, the solvent is evaporated, the residue is dissolved in MeOH/toluene (1:9, ~30ml), filtered from some insoluble material and evaporated *in vacuo*. The crude residue is purified by preparative HPLC (Nomura Chemical RPFullerene, with toluene/MeOH 1:9 as eluant) to give on evaporation the desired fullerene as an orange powder (99.5mg, 0.064mmol, 12% yield). Its <sup>1</sup>H NMR (400 MHz, acetone-*d*<sub>6</sub>, TMS) has δ at 7.32-7.28 (m, 8H), 7.23 (d, *J* = 8.0Hz, 2H), 6.69-6.64 (m, 10H), 4.55 (s, 1H), 2.38 (s, 2H), 2.31 (d, *J* = 14.4Hz, 2H), 2.23 (d, *J* = 14.4Hz, 2H), 2.11 (d, *J* = 14.8Hz, 2H), 2.05 (d, *J* = 14.8Hz, 2H), 0.36 (s, 6H), 0.35 (s, 6H), 0.34 (s, 6H), 0.33 (s, 6H), 0.27 (s, 6H); and the <sup>13</sup>C NMR (100 MHz, acetone-*d*<sub>6</sub>) has δ at 159.84, 159.81, 158.36, 155.22, 155.18, 154.80, 150.53, 149.13, 149.19, 149.05, 148.74, 148.66, 148.41, 148.18, 148.70, 147.66, 147.30, 146.97, 146.52, 146.22, 145.93, 145.24, 144.07, 143.53, 143.49, 142.76, 136.07 (C<sub>6</sub>H<sub>4</sub>), 136.03 (C<sub>6</sub>H<sub>4</sub>), 127.99 (C<sub>6</sub>H<sub>4</sub>), 127.84 (C<sub>6</sub>H<sub>4</sub>), 118.89 (C<sub>6</sub>H<sub>4</sub>), 115.92 (C<sub>6</sub>H<sub>4</sub>), 115.87 (C<sub>6</sub>H<sub>4</sub>), 115.82 (C<sub>6</sub>H<sub>4</sub>), 55.66 (CCH<sub>2</sub>SiMe<sub>2</sub>), 54.11 (CCH<sub>2</sub>SiMe<sub>2</sub>), 53.69 (CCH<sub>2</sub>SiMe<sub>2</sub>), 31.20 (C<sub>60</sub>CCH<sub>2</sub>Si), 30.39 (C<sub>60</sub>CCH<sub>2</sub>Si), 30.35 (C<sub>60</sub>CCH<sub>2</sub>Si), -0.22 (SiMe<sub>2</sub>), -0.30 (SiMe<sub>2</sub>), -0.34 (SiMe<sub>2</sub>), -0.47 (SiMe<sub>2</sub>); the HRMS (APCI+) has calculated for C<sub>105</sub>H<sub>67</sub>Si<sub>5</sub>O<sub>5</sub> (M-2H)<sup>+</sup> 1547.3835, and the found mass is 1547.3840.

This fullerene now has five phenolic groups for condensation with lipid molecules. When a mixture of it (9.40mg, 6.07 $\mu$ mol) and 4-(dodecan-1-yloxy)benzoyl chloride (13.0mg, 9.81 $\mu$ mol) in THF (3.0ml) is treated with Et<sub>3</sub>N (5.8  $\mu$ L, 42mmol) and 4-dimethylaminopyridine (3.5mg, 29 $\mu$ mol), stirred at 25° for 3 hours, evaporated *in vacuo*, and the residue in toluene is filtered through a silica gel pad, and then purified by GPC (JAIGEL 3H, eluting with toluene), precipitated with MeOH/toluene as above and dried *in vacuo* (at 25° for 12 hours), it provided **1,4,11,15,30-pentakis{4-[4-(dodecan-1-yloxy)benzoyloxy]phenyldimethylsilylmethyl}-2H-1,2,4,11,15,30-hexahydro-(60)fullerene** (12.9mg, 71% yield). Similarly the [4-tetradecan-1-yloxy] (75% yield), [4-hexadecan-1-yloxy] (72% yield), [4-octadecan-1-yloxy] (61% yield) and [3,4-di(octadecan-1-yloxy)] (37% yield) *derivatives* were obtained in high purity and characterised by <sup>1</sup>H NMR, <sup>13</sup>C NMR and ACPI MS analysis. These form beautifully coloured liquid crystals and X-ray diffraction studies showed interesting stacking of these fullerene molecules. [Matsuo, Muramatsu, Hamasaki, Mizoshita, Kato and Nakamura *J Am Chem Soc* **126** 432 2004.]

[Further reading: H.W. Kroto et al. *Chem Rev* **91** 1213 1991; H. Kroto, Fischer and Cox *Fullerenes*, Pergamon Press, Oxford, 1993 ISBN 0080421520; Kadish and Ruoff (Eds) *Fullerenes: Recent Advances in the Chemistry and Physics of Fullerenes and Related Materials*, The Electrochemical Soc. Inc, Pennington, NJ, 1994, ISBN 1566770823; Smalley *Acc Chem Res* **25** 98 1992, and following papers; Hammond & Kuck *Fullerenes: Synthesis, Properties and Chemistry of Large Carbon Clusters*, American Chemical Society, Washington, 1992, ISBN 0-841221820; K. Jino *Separation of Fullerenes by LC*, RSC Publ., 1999, ISBN 9780854045204; S. Nagase & T. Akasaki *Endofullerenes: a new family of carbon clusters*, Springer, 2003, ISBN 9781402009822; F.J.M. Reitmeijer *Natural Fullerenes and related structures of elemental carbon*, NetLibrary, 2006, eBook ID 190026, eISBN 9781402041358; P.W. Fowler & D.E. Manopoulos *An Atlas of Fullerenes*, Dover Publications Inc, 2007 ISBN 9780486453620; P. O'Brien, H. Criaghead, H.W. Kroto, F. Langa and J-F. Nierengarten *Fullerenes: principles and applications* RSC Publ., 2007, ISBN 9780854045518; *Fullerenes, Nanotubes and Carbon Nanostructures*, Marcel Dekker Inc, New York, on line series on *World Wide Web.online*; N. Chaniotakis *Fullerenes-bifunctionalisation (nanostructured for biosensing)* in "Nanomaterials for Biosensors", C. Kumar ed., Wiley-VCH, 2007, ISBN 9783527313884; C.N. Kramer *Fullerene research advances*, Nova Science Publishers Inc, 2007, ISBN 9781600218248; M. Lang *Progress in fullerene research*, Nova Science Publishers Inc, 2007, ISBN 9781600218415; S. Margadonna *Fullerene-related materials: recent advances in their chemistry and physics*, Springer 2007, ISBN 9781402044588.]

## CARBON NANOTUBES (CNTs) AND RELATED MATERIALS

The development of carbon nanotubes (CNTs) since their discovery by Sumio Iijima [*Nature* **354** 56 1991] has been explosive, not only because of their mechanical, thermal and electronic properties but because they can be modified physically and chemically in a variety of ways to suit a plethora of applications. Structurally, they are made of rolled up sheets of carbon atoms forming fused hexagonal carbon rings just like a monolayer of graphite. The sides of the sheet are fused to form complete tubes. When one or both ends of a tube come together, like the bottom of a test tube, the round portion is more like half a 'fullerene ball'. The way that the carbon atoms in the half fullerene ball 'finger tip' structure are stable is by forming both hexagonal and pentagonal rings (see Marchand et al. *Nano Lett* **9** 2961 2009, FEM imaging below).

Many carbon nanotubes are now available commercially in chemical catalogues, and a selected number are entered here. Their chemical and biochemical applications include diagnostics, novel devices for rapid DNA sequencing, human personalised medicine, bio-labeling for cancer research, novel biomaterials for human implants, and for tissue engineering implants. The non-chemical applications have also been extended to attachments to surfaces with patterned and unpatterned textures to form conductors, and exciting applications as for nano-filters, in LEDs, nanostructured optical fibres, lasers, and solar cells. Carbon nanotubes are very strong materials and could be ~200 stronger than steel. When mixed with building materials, such as cement or aluminium composites, they not only increase their strength but make them much lighter. Carbon sheets such as graphene are conducting surfaces and have found extensive use in touch screens, i-Pods, mobile phones to name a few. [Haddon *Carbon Nanotubes (editorial)*, *Acc Chem Res* **35** 997 2002, see also P. Calvert *A recipe for strength*, *Nature* **399** 210 1999].

### PREPARATION OF CNTs (NTs)

**MWCNTs** (multi walled CNTs) were first prepared by the arc discharged evaporation method (as for fullerene synthesis) where the needles (nanotubes) grew at the negative end of the carbon electrode in an argon filled vessel (at 100Torr). TEM micrographs showed tubes of 5 concentric graphic sheets (6.7nm diameter), 2 concentric graphic sheets (5.5nm diameter, DWCNT, double walled CNTs), and 7 concentric sheets (2.2nm diameter MWCNT). Electron diffraction patterns show that the tubes were formed from rolled up graphite sheets of carbon hexagons in helical structures. [Iijima *Nature* **354** 56 1991.] In chemical arc deposition, acetylene (at partial pressure  $\sim 10^{-2}$ mbar, 0.0075mm Hg at 0°) is often used as the hydrocarbon gas. J-M Bonard and coworkers prepared CNTs by growing them on 0.3mm diameter wires of Kanthal (Fe-Al-Cr alloy) in the centre of the anode at the rather low temperatures of 700-900°, hence the name *Cold Atmosphere Chemical Vapour Deposition CVD* (CACVD, compare below). The technique produces homogeneous films of well-graphitised NTs allowing control of their length and density. [Crocetti et al. *Chem Vap Deposition* **8** 89 2002, Bonard et al. *Appl Phys Lett* **81** 2836 2002]. MWCNTs are generally formed to a smaller or larger extent in the procedures used together with small quantities of buckyballs and amorphous carbon which can be invariably sublimed out, annealed out, or separated by chromatography. Annealing at high temperatures ( $\sim 2500$ - $2700^\circ$ ) removes volatile carbon, metals (generally bound to the open end of the tubes) and improves the structural order of the tubes.

**SWCNTs** (single walled CNTs): Among the techniques that have been developed for making SWCNTs are EA (electric arc) methods which use a carbon source and catalytic amounts ( $\sim 1$  to 6 atom%) of transition metals, commonly Ni/Co, Ni/Y, Rh/Pt or Fe that are allowed to supersaturate the carbon [Kokai et al. *J Phys Chem B* **103** 4346 1999, Sen et al. *Chem Phys Lett* **349** 383 2001]. The methods have been studied in detail in order to optimise the sizes and yields of CNTs. The technique developed by R.C Haddon and co-workers [Itkiss et al. *J Phys Chem B* **108** 12770 2004], and the purity estimated by solution phase NIR [Itkiss et al. *Nano Lett* **3** 309 2003] adopts the EA method where an electric arc discharge is applied between a graphite cathode and composite anode containing catalytic amounts of metal(s), e.g. 4:1 atom% of Ni/Y, under He buffer gas at  $\sim 680$ Torr, and arc current 90A with a 10mm arc gap. These conditions produce high yields of AP-SWCNT (high purity, As-Prepared nanotubes). The purity can be assessed by Near-Infrared (NIR) spectroscopy. Large scale production of SWCNTs by the EA technique is achieved by using an arc generated between a graphite rod (16mm diameter, 40mm long) cathode and an anode of graphite (6mm diameter, 100mm long) in which a hole 3.5mm diameter, 40mm deep, is drilled and filled with a mixture of a metallic catalyst (Ni 4.2 atom%/Y1 atom% and graphite) in a He atmosphere (660mbar) formed by a current of 100A/30V and keeping a constant distance ( $\sim 3$ mm) between the electrodes. Within  $\sim 2$  minutes, the total carbon mass was in  $\sim 2$ g quantities consisting of  $\sim 70\%$  SWCNTs in highly crystalline bundles. The nature of the product(s) was assessed by SEM, high resolution TEM, Raman spectroscopy and X-ray diffraction. [Journet et al. *Nature* **388** 756 1997.]

Another technique developed by R.E. Smalley and coworkers is the laser ablation of carbon targets (*Carbon Vapour Deposition*, CVD), e.g. laser-vaporised carbon-Ni/Co mixture (e.g. 1:1 1.2 atom% in C) at high temperatures (e.g. 1200°C) and vacuum (e.g. 500Torr) and an argon stream. [Thess et al. *Science* **273** 483 1996, Sen et al. *Chem Phys Lett* **332** 467 2000, Koran et al. *J Vac Sci Technol A* **13** 1171 1995]. This CVD produces high yields ( $\sim 70\%$ ) of SWCNTs as assessed by X-ray diffraction and electron microscopy (EM) methods at a rate of  $\sim 80$ mg/day. The production of SWCNTs can be increased to  $\sim 1$ g/day by using a 2" tube and dual laser pulses. However, by modifying the apparatus, using up and down scanning of the laser pulses onto a rotating carbon-Ni/Co target in a 4" diameter tube, a lower temperature (1100°C) can be used to generate 20g of 40-50vol% SWCNT material in 48 hours of continuous (largely unattended) operation [Rinzler et al. *Appl Phys A* **67** 29 1998]. The high purity was checked by EM, XRD, Raman spectroscopy and TGA. Note that higher temperature operations produce tubes of larger diameters, albeit well within the nanometer scale.

In a more recent, essentially CVD procedure, S.T. Purcell and co-workers [Marchand et al. *Nano Lett* **9** 2961 2009 which provide an animated supplement, and summarised by H. Birch 'Nanotube growth on camera', *RSC Chemistry World* **6**(9) 26 2009] have demonstrated by *Field Emission Microscopy* (FEM) how the growth of SWCNTs takes place. They used a W-tip connected to the electrodes (at 1200°C) in a chamber at ultra high vacuum ( $5 \times 10^{-10}$  Torr base pressure). The W-tip (tip radius  $\sim 60$ nm) is first covered with a graphite diffusion barrier by heating in acetylene (at  $P_{C_2H_2} \sim 1 \times 10^{-4}$  Torr), and Ni nanoparticles are induced to grow directly onto the tip [by dewetting, from a circular Ni wire in the chamber using CVD in  $C_2H_2$  ( $\sim 1-2 \times 10^{-7}$  Torr during growth) at 800°C] while FEM imaging is performed. The images are collected on camera, and a frame-by-frame analysis of the video showed that the SWCNT grows at one end of the carbon tube at a time. The tube

rotates in discrete steps, ~24 per rotation (half of the number of carbon atoms on the circumference of a common SWCNT) with the CNT turning ~180 times during its 11 minutes growth (as in the 'screw-dislocation-like' model). Growth starts at the Ni atom (of the catalyst), the carbon attaches itself to the metal then it forms carbon rings (as hexagons, because these are the more stable configuration), and the molecule rotates as the tube is being formed. In the end, the tube is sealed by forming a semi-fullerene ball 'finger tip' structure, and produces a 'test-tube like' nano carbon structure with the metal atom at its mouth. The metal can then be removed by heating the SWCNT strongly (annealing).

The *HiPco* process for preparing SWCNTs developed by R.E. Smalley and co-workers [Nikolaev et al. *Chem Phys Lett* **313** 91 1999] involves formation of the catalysts by *in situ* thermal decomposition of Fe(CO)<sub>5</sub> in a heated flow of CO in the gas phase at 1-10atm and 800-1200°C. By adjusting the processing parameters (best at 1-10atm and 800-1200°C), SWCNTs were produced in high yields (79 mole%, i.e. 44 wt%) with narrowest tubes (as small as 0.7nm in diameter, same as in C<sub>60</sub> fullerene) at the rate of 1.24mg/hour. The structures were confirmed by TEM, SEM, EDX and TGA. An important advantage of this procedure is that it is a continuous-flow process which can be used for bulk production of narrow SWCNTs. See below for purification of HiPco tubes.

### PURIFICATION OF CNTs (NTs)

**MWCNTs:** Raw tubes, as prepared by the EA method, contain about one third of other nanomaterial. This material is best removed by oxidation. Pure tubes can thus be prepared, although the procedure is wasteful. The purest tubes are obtained when 99% is oxidised; but when 95% is oxidised only 10-20% of product contains pure nanotubes. A ground raw sample is placed in an oven and the temperature is raised to 750° in air or oxygen for ~30 minutes until ~1% is left. This consists mostly of pure tubes with length/diameter ratios that exceed 100, and traces of open cylinders with the ratio of 20. [Ebbesen et al. *Nature* **367** 519 1994, see also Chen et al. *Adv Mater* **8** 1012 1996.]

MWCNTs, prepared by the CVD procedure using ferrocene and xylene (0.75% Fe/C) as catalyst and precursor with argon/H<sub>2</sub> at ~700° [Andrews et al. *Chem Phys Lett* **303** 467 1999 (a step closer to commercialisation), Rao et al. *Appl Phys Lett* **76** 3813 2000], or acetylene as precursor and 2.5wt% Co-2.5wt% Fe/NaY zeolite as catalyst with argon/H<sub>2</sub> at ~700° [Bulusheva et al. *J Phys Chem B* **105** 4853 2001], were purified by heating the raw tubes (200mg) in 2.6M aqueous HNO<sub>3</sub> (40ml) for 48 hours. The cooled solution was centrifuged (use PTFE tubes, ~20,000g/30 minutes), the sediment was washed/re-centrifuged with de-ionised H<sub>2</sub>O until the supernatant was barely acidic and the MWCNT sediment was dried *in vacuo*. This material was satisfactory for functionalising [Lin et al. *J Phys Chem B* **106** 1294 2002]. Alternatively, to a dispersed suspension of MWCNTs (100mg), formed by sonication (1 minute), is added an acidic solution of KMnO<sub>4</sub> (4g in 120ml of 1N H<sub>2</sub>SO<sub>4</sub>) dropwise and refluxed for 12 hours, cooled, and centrifuged. The sediment is treated with concentrated HCl (20ml) and refluxed for 24 hours to dissolve the MnO<sub>2</sub>, cooled, centrifuged and the sediment is refluxed with 68% (azeotropic) HNO<sub>3</sub> (40ml) for 24 hours, cooled, centrifuged again, and the solids are washed repeatedly (using several centrifugation/washing cycles) with de-ionised H<sub>2</sub>O until almost neutral then dried in a vacuum oven. This procedure exposes the CO<sub>2</sub>H groups on the CNTs and can be used for conversion to COCl groups, and further reactions to make soluble MWCNTs. [Fu et al. *Nano Lett* **1** 439 2001.]

**SWCNTs: (i)** Mechanically ground cloth-like raw SWCNTs from the electric arc process (EA) can be purified in reproducibly high yields (optimally 25-30wt% and <1wt% of transition metal) in two steps. The first involves thermal annealing in air by rotating the powders at 470°C for 50 minutes in a quartz tube at 30rpm which burns the carbonaceous particles out, and the second is acid treatment which etches out the metals. The powder is immersed in aqueous 6M HCl for 24 hours, filtered (or centrifuged) several times until the colour of the acid is unaltered, and then washed with de-ionised H<sub>2</sub>O. The SWCNTs are then 'unbundled' by boiling in 30% aqueous HNO<sub>3</sub> for 4-6 hours, and the suspension is filtered with a polytetrafluoroethylene (PTFE) membrane in deionised H<sub>2</sub>O, rinsed and dried to give a grayish black, thin mat of SWCNTs. The purity at each stage can be observed by SEM, TEM, FT-Raman spectroscopy and TGA which usually show ~96% purity after HCl treatment. The HNO<sub>3</sub> treatment, however, appears to break down the CNTs into small pieces and sometimes forms MWCNTs. [Moon et al. *J Phys Chem B* **105** 5677 2001.]

**(ii)** A gas-phase purification procedure developed by Margrave and co-workers [Zimmerman et al. *Chem Mater* **12** 1361 2000] involved purging the CNT sample [prepared by the pulsed laser method (Rinzler et al. *Appl Phys A* **67** 29 1998), from Graphite (1-2mm), or from EA] in a quartz tube with a mixture of Cl<sub>2</sub> (7.2ml/min), H<sub>2</sub> (2.7ml/min) and argon (3.0ml/min) for 1 hour at ~25°, which is then lowered into a furnace at 500° while the evolving gases (CO, COCl<sub>2</sub>, CCl<sub>4</sub> and CO<sub>2</sub>) are monitored by IR. Purging is continued until the CO partial

pressure is  $<0.5$ Torr. The temperature of the sample is lowered to  $\sim 25^\circ$ , sonicated in DMF/0.6M HCl (1:1) to remove metals, then pure DMF, filtering, and washing with MeOH to form an SWCNT 'paper'. (*Alternatively*, the metals, if present, can be removed by sublimation in HCl gas at higher temperature.) The SWCNT 'paper' is dried at  $160^\circ$  and the purity assessed by SEM, TGA and UV-VIS. Any fullerenes present may be removed by a toluene reflux or by sonication. The yield is  $\sim 15$ wt% to give CNTs that are generally more reactive due to their larger curvature.

(iii) Raw HiPco SWCNTs can be purified in a multi-step process. Low density raw HiPco ( $\sim 100$ mg,  $d \sim 0.01$ g/cc) is compressed onto dry filter paper (vacuum), placed in a ceramic boat and inserted into a quartz tube furnace while a gas mixture of 20%  $O_2$  (or air) in argon (wetted by bubbling through  $H_2O$ ) is allowed to flow over the sample in the furnace at 100cc/sec. The nanotubes are heated at  $225^\circ$  for 18 hours, sonicated for  $\sim 15$  minutes (or stirring overnight) in concentrated HCl (yellow colour due to  $Fe^{3+}$  ions). The suspension is filtered onto a Teflon membrane (47mm,  $1.0\mu m$  pore size), washed several times with deionised  $H_2O$ , MeOH, and dried in a vacuum oven at  $100^\circ$  for 2 hours (weight loss 33.7%; 0.67% residual metal). This oxidation with wet  $O_2$  (or air) in argon, and extraction cycle is repeated at  $325^\circ$  for 1.5 hours (further weight loss 8.3%, 0.05% residual metal), and again at  $425^\circ$  for 1 hour (further weight loss 22.9%, 0.09% residual metal). After drying *in vacuo*, the HiPco tubes are annealed at  $800^\circ$  in argon for 1 hour (further weight loss 4.2%, 0.03% residual metal). Assessment of purity is by SEM, TEM, TGA, Raman and UV-VIS-NIR spectroscopy. The onset of oxidation of the smaller diameter HiPco SWCNTs is *ca*  $100^\circ$  lower than for the purification of large diameter tubes obtained in the laser-oven process, and is consistent with the greater steric strain present in small diameter SWCNTs. [Chiand et al. *J Phys Chem B* **105** 8297 2001.]

(iv) A general purification procedure for raw SWCNTs obtained by the EA process (reaching 2-3g, containing  $\sim 70\%$  SWCNTs) [Ebbesen & Ajayan *Nature* **358** 220 1992, Journet et al. *Nature* **388** 756 1996], or the laser ablation method (reaching  $\sim 20$ g, containing up to 50% SWCNTs) was developed by A. Hirsch and coworkers [Holzinger et al. *Appl Physics A* **70** 599 2000]. The process involves three steps. (a) Raw material (100mg) was refluxed with 65%  $HNO_3$  (150ml) for 3 hours, and centrifuged. The sediment was washed with  $H_2O$  and centrifuged repeatedly until the supernatant (first colourless then darkens, still contains some SWCNTs and was kept for further use) has pH  $\sim 1$ . The sediment is suspended in distilled  $H_2O$ , subjected to two and three 0.5 seconds of ultrasonic pulses when the liquid became weakly acidic (pH 3-4), almost free from metal catalyst, and contained  $\sim 1$ mg/ml  $H_2O$  of only SWCNTs and nanoparticles. For storage, the pH is adjusted to 8-9 with  $K_2CO_3$  which increased the repulsive interactions between the carboxylate groups of the SWCNTs. (b) The suspension is sonicated for 1 minute which reduces the nanoparticles to smaller pieces. The tubes and bundles were also smashed and the difference in size between the SWCNTs and degraded particles increases (as observed by AFM). (c) This most important step involved elution of the suspension through a column of potassium polyacrylate which was swollen (the cross-linked polymer particle size swells from  $<1\mu m$  to  $\sim 300\mu m$ ) creating pores large enough to entrap nanoparticles as well as SWCNTs and bundles of comparable size. [Optimum: 15ml of swollen polymer per 1ml of a 1g/L aqueous SWCNTs, and choosing a diameter to give a column height of 6-7cm held by a glass filter pore size 2]. A vacuum is applied at the bottom of the column which squeezes the swollen polymer particles. The SWCNTs are so large that they cannot now be entrapped and elute freely to give large quantities ( $> 40\%$  of the total mass) in the first fraction. The remaining material contains a large portion of SWCNT fraction of lesser quality which can be recycled. Most of the degraded material and smaller nanoparticles remain in the cavities. The purification can be followed by AFM and the purity can be assessed by Raman spectroscopy. In the Raman spectrum, *pristine* SWCNTs have a doublet at 1546 and 1575  $cm^{-1}$  (of about equal intensity — G-lines) and a less intense band (D-line, *ca*  $1/4.75$  th of the G-lines) at  $\sim 1342$   $cm^{-1}$ . The relative intensity of the G-lines increase as the purity increases. The broadness of the D-band may originate from disordered or carboxylated carbon formed by oxidation.

#### PREPARATION AND PURIFICATION OF FULLERENE PIPES (TUBES)

These are short SWCNTs of 100 to 300nm lengths and  $\sim 1.2$  to 1.4nm in diameter which are most useful for preparing functionalised SWCNTs. They are distinct from the above *fullerenes* which are "bucky balls" and are made up of tubes consisting of a network of hexagonal carbon rings like in CNTs. As prepared, the CNTs are long almost endless single walled highly convoluted, entangled, bundled tubes. R.E. Smalley and coworkers [Rinzler et al. *Appl Phys A* **67** 29 1998, Liu et al. *Science* **280** 1253 1998] have purified, and cut such material from large batches into small lengths of open-ended tubes, *Fullerene pipes*, which form stable colloidal suspensions in the presence of detergents. These can be made to react like ordinary organic molecules to form any variety of derivatives. In a typical preparation, the raw tubes (10g) in 2.6M  $HNO_3$  were refluxed for 45

hours, cooled, and the black solution was centrifuged in PTFE (Teflon) tubes (20,000g, 20 minutes, Sorvall R5C centrifuge). The supernatant was decanted off, the sediment was resuspended in de-ionised H<sub>2</sub>O by vigorous shaking and the centrifuging/washing cycle repeated 3 to 4 times until the supernatant was almost neutral, although black in colour. The acid treatment not only assists in separating the bundles, but also oxidises the material, thus producing carboxylic acid groups at the ends as well as at the sides of the tubes. This makes 'holes' in the sides of the tubes and causes them to be 'cut' [cf K. Kinoshita *Carbon Electrochemical and Physicochemical Properties* Wiley, New York 1988, ISBN 978047184802-8]. The sediment as examined by SEM and TEM contains carboxylated carbon tubes and other species such as fullerenes, polycyclic aromatic sheets and cross-linked sheets edged with CO<sub>2</sub>H groups which, by virtue of de-protonation allow mutual repulsion between the particles. The sediment (using 10g) was dispersed in aqueous NaOH (1.8L, pH 10) containing 0.5vol% of the non-ionic surfactant Triton X 100 by sonication for ~1 hour, and purified by *cross-flow-filtration* (CFF, cellulose ester M22M 600 01N, mini-Kros® Spectrum cartridge: 0.6mm diameter, 200nm pores and 0.56m<sup>2</sup> of surface area) against a buffer of aqueous NaOH (40L, pH 10) containing 0.2vol% of Triton X 100. At a head pressure of 5-6psi and flow rate of ~70ml/min the process was complete in ~10 hours. Salt was removed by washing with de-ionised H<sub>2</sub>O (10L). Clogging occasionally occurred and was overcome by momentarily reversing the flow before continuing. Clamping off the buffer line and opening a vent allowed concentration of the SWCNTs down to ~200ml of solution. When a portion of this solution was filtered through a PTFE membrane (Millipore LS 5µ pore) using a vacuum, it deposited a thick pad of SWCNTs on the membrane. After washing with MeOH, the pad can be easily peeled off to give a freestanding mat which was called 'Bucky Paper' and was estimated as typically 10 to 20% yield depending on the initial raw material. TEM images of the SWCNT 'paper' showed that the pipes still contained impurities which were removed by two further oxidising acid treatment which 'cut' (by reacting at the sides of the tubes) and 'polished' the tubes.

In a typical *cutting*, the purified SWCNT 'Bucky Paper' (10mg) was suspended in a 3:1 mixture of concentrated H<sub>2</sub>SO<sub>4</sub>/HNO<sub>3</sub> (40ml), sonicated (55kHz) at 35° to 40° for 24 hours. The suspension was diluted with H<sub>2</sub>O (200ml), larger tubes were collected onto a 100nm pore filter membrane (Millipore, VCTP type), and washed with 10mM NaOH solution. *Polishing* was performed by suspension in a 4:1 mixture of concentrated H<sub>2</sub>SO<sub>4</sub>/30% H<sub>2</sub>O<sub>2</sub>, stirring at 70° for 30 minutes, filtering and washing again on a 100nm filter, and re-suspending at 0.1mg/ml of 0.5wt% Triton X100 in H<sub>2</sub>O to avoid flocculation. On a larger scale CFF purification was carried out after each oxidation step. Finally, to obtain best purified 'Bucky Paper', annealing *in vacuo* at 1200° was necessary. Note that in 3:1 H<sub>2</sub>SO<sub>4</sub>/HNO<sub>3</sub> at 70° the average cutting was at a rate of ~130nm/hour, and in 4:1 H<sub>2</sub>SO<sub>4</sub>/30% H<sub>2</sub>O<sub>2</sub> the shortening rate was at ~200nm/hour.

*Characterisation* of these fullerene tubes was done by electro-deposition [developed by Smalley and coworkers, see above references] which drives the suspended tubes onto the surface of highly oriented pyrolytic graphite surface and are scanned by AFM. The molecular nature of the tubes was demonstrated by converting their CO<sub>2</sub>H groups to COCl, followed by reaction with H<sub>2</sub>N(CH<sub>2</sub>)<sub>11</sub>SH and exposure to 10nm gold particles. Gold (which reacts with SH) was only found on the carboxylated fullerene tubes.

### SOLUBILISATION OF CNTs

CNTs, whether MWCNTs or SWCNTs, readily form bundles, and attempts to separate the bundles by means such as sonication or use of detergents have been uniformly unattractive. The carbon material is 'non-wettable' and quite insoluble in polar and non-polar liquids. Considerable effort has been expended to find means of separating the bundles and much success has been achieved by functionalising the carbon matrices, which destabilise the interaction between the fibres, but also introduce bound groups which afford solubility in organic solvents as well as aqueous solutions. These provide opportunity for purification because the partly purified CNTs are not homogeneous. The non-homogeneity is not only in the length and/or diameter, but also in their nature since during the formation of the tubes, the row of hexagonal rings do not only align to form straight tubes (with parallel hexagonal rows), but can also form spiral tubes in which the row can be out of line by one and/or more than one hexagon while forming the tube. [For the *Dislocation theory of chirality-controlled nanotube growth* see B.I. Yakobsen and coworkers *Proc Natl Acad Sci USA* **106**(8) 2506 2009, and for a review see H. Birch *RSC Chemistry World* **6**(3) 25 2009.] This provides the opportunity for a variety of chiral CNTs. In his respect M Zheng and coworkers [Zheng & Semke *J Am Chem Soc* **129** 6084 2007, Tu et al. *Nature* **460** 250 2009, see also P. Broadwith *RSC Chemistry World* **6**(8) 28 2009] have selected short strands of DNA [8 to 12 mers, with repeating nucleotides e.g. (GTC)<sub>2</sub>, (GTC)<sub>2</sub>GT, (GTT)<sub>3</sub>G, (TAT)<sub>4</sub> etc] which entwine around the SWCNTs (e.g. produced from HiPco, see above) of *ca* 1nm diameter with different selectivities depending on the '*n,m* chirality' of the nanotubes [e.g. (8,6), (9,4), (7,6), (6,5) etc respectively] to form barrels around the

tubes. The mixture is dispersed in appropriate aqueous solutions (e.g. 0.1M NaCl, 0.1M NaCl/10% glycerol) by sonication (~60 minutes), incubated for a period of time, centrifuged (16,000g/90 minutes) and the supernatant is fractionated by Size Exclusion Chromatography (SEC) through an IEX resin (CNT-NS 1500, Biochrom, Terre Haute, IN, USA), and the various DNA-wrapped-SWCNTs are eluted with 2x(0.3M NaCl, 0.03M sodium citrate)/0.5mM EDTA/pH 7.0 and a 0 to 1M sodium benzoate gradient. The fractions are monitored at 300nm and 425nm. The DNA interaction is electrostatic and is easily removed by thermal treatment.

Strong acids solubilise SWCNT and MWCNT bundles by protonating them causing the strands to separate, but chlorosulfonic acid is by far (~1,000 times) the best acid as shown by M. Pasquali and the late R.E. Smalley [Davis et al. *Nature Nanotechnology* **4** 830 2009, summarised by P. Broadwith RSC *Chemistry World* **6**(12) 32 2009]. SWCNTs spontaneously dissolve in ClSO<sub>3</sub>H at a concentration of as low as 0.5wt%. The Phase Diagram in strong H<sub>2</sub>SO<sub>4</sub> exhibited true thermodynamic behaviour forming isotropic (I), liquid crystal (LC), crystal solvate (CS) and solid (S) phases at different concentrations of SWCNTs, ClSO<sub>3</sub>H and H<sub>2</sub>SO<sub>4</sub>. When an 8.5vol% of SWCNTs was coagulated in 96% aqueous H<sub>2</sub>SO<sub>4</sub> by extrusion through an orifice while co-flowing with CHCl<sub>3</sub> or CH<sub>2</sub>Cl<sub>2</sub>, the resulting nanotube bundles (10-100nm) combined into larger bundles to form macroscopic fibres which were easily spun into tens of meters of continuous fibres in a matter of minutes. These fibres can yield a strength in excess of 320 MPa. By sandwiching the dope between glass plates, long ribbons of material were produced, and possibly any shapes (e.g. coils) can be made. The technological applications are enormous.

For examples of some currently commercial CNTs and functionalised CNTs see the following paragraphs.

### CARBON NANOTUBES (CNTs)

**Carbon annealed nanodiscs/nanocones** [7440-44-0]. These are available as a black powder containing a mixture of 20% cones, 70% discs and 10% of carbon black impurities by weight. The carbon cones have a nearly perfect geometry. The five theoretically possible cones with apical angles of 19.2°, 38.9°, 60°, 84.6° and 112.9° are all present in the mixture with cone lengths of 0.3-0.8μ, maximum base diameter of 1-2μ and wall thickness of 20-50 nanometers. The diameter and thickness of the discs are 0.88-3.5μ and 20-50 nanometers respectively. A product of almost 100% carbon is obtained by annealing at 2500-2700° which increases the structural order and reduces the concentration of impurities particularly of metals (from catalyst). They are similar to multi-walled carbon nanotubes (MWCNTs).

**Carbon multi-walled nanotubes (MWCNTs or MWNTs)** [1333-86-4]. These are produced as a black powder by chemical vapour deposition (CVD) and contain >90% of nanotubes which are almost free from metal impurities. The tubes are ~7μ in length and ~140 nanometers in diameter. Multi-walled smaller nanotubes of ~1.5μ in length and ~20-25 nanometers in diameters are similarly prepared. When the multi-walled carbon nanotubes are 'arc-produced' they are even smaller in size with ~2μ in length and ~2-50 nanometers in diameter [308068-56-6]. This 'arc process' produces a mixture of 55-65wt% of straight multi-walled nanotubes and 35-45wt% of graphite nanoparticles. They contain almost 100% carbon. By growing the tubes at 3000-4000° using this process they contain less defects than by other procedures, and are purified in this way. They are stable in air up to 700°.

MWCNT arrays for nanoelectronics with nickel-carbide particle tips with same diameters as above are also available. Vertically aligned MWCNTs on a silicon wafer substrate with low resistivity (1-30 ohm-cm), as well as on a copper wafer substrate (1cm x 1cm x 0.05cm high conductivity low-oxygen copper) have been prepared. The latter are produced by plasma-enhanced chemical vapour deposition (PECVD). MWCNTs have also been made in 3-19, 7-13 and 5-15 graphene layers wall thicknesses. See below for breaking up the bundles and graphene.

**Carbon double-walled nanotubes (DWCNTs or DWNTs).** These have been prepared by the *Carbon Vapour Deposition (CVD)* method with >90% carbon content and 50-80% DWCNT. They can be prepared with approximately the following dimension 5nm OD, 1.3-2.0nm ID and 50μm length; and with 0.12-0.14g/ml bulk density and >600m<sup>2</sup>/g surface area. The rest (10-40%) consists of amorphous carbon and residual metal catalysts which are deliberately placed in the original carbon in order to synthesise the carbon tubes.

**Carbon single-walled nanotubes (SWCNTs or SWNTs)** [308068-56-6] **M 3652-3607, d<sup>25</sup> 1.7-1.9g/ml.** The SWCNTs are of particular use for derivatisation and the applications stated above. They can be prepared in various bundle dimensions, viz:  $d = 1.2-1.5\text{nm}$ ,  $l = 2-5\mu\text{m}$  (40-60% SWCNT, arc method);  $d = 1.1-0.5\text{nm}$ ,  $l =$

100 $\mu\text{m}$  (>50% SWCNT, Carbon Vapour Deposition, CVD, method);  $d = 2\text{-}10\text{nm}$ ,  $l = 1\text{-}5\mu\text{m}$  (50-70% SWCNT, arc method); individual short tubes  $d = 1\text{-}2\text{nm}$ ,  $l = 0.5\text{-}2\mu\text{m}$  (90% SWCNT, electric arc, EA, method), and individual long tubes  $d = 1.3\text{-}1.5\text{nm}$ ,  $l = 1\text{-}5\mu\text{m}$  (40-60% SWCNT, arc method, EA) and the HiPco process (see above).

### FUNCTIONALISED CNTs.

MWCNTs and SWCNTs have been successfully functionalised, and the products have found numerous applications. They are conducive to further chemical reactions as they are readily soluble, thus amenable to solution chemistry [see R.C. Haddon and coworkers in *Chemistry of Single-Walled Carbon Nanotubes*, Niyogi et al. *Acc Chem Res* **35** 1105 2002; and Sun and coworkers in *Functionalized Carbon Nanotubes: Properties and Chemistry*, Sun et al. *Acc Chem Res* **35** 1096 2002]. When purified by the acid and oxidising conditions the tubes invariably are oxidised to form CO<sub>2</sub>H or quinone groups, particularly at the ends of the tubes. Depending on the conditions, ‘cutting’ may occur leaving open-ended tubes (see fullerene tubes above). Annealing at high temperatures *in vacuo* causes some decarboxylation and closing up of the ends which form half fullerene tips (like the end of a glass test tube, see above). Because of the larger curvature at the tips than on the sides of the tubes, the tips are more readily functionalised (due to the easier conversion of sp<sup>2</sup> to sp<sup>3</sup> carbon atoms); but then again they would lose the added function more readily on annealing. The tips are also more easily cleaved by oxidants (KMnO<sub>4</sub>, OsO<sub>4</sub> and RuO<sub>4</sub>) at 100° under acidic conditions [Hwang *Chem Commun* 173 1995]. Functionalised SWCNTs dissolve in organic solvents (e.g. THF, CH<sub>2</sub>Cl<sub>2</sub>) and can be examined by optical spectroscopy. Tubes with CO<sub>2</sub>H groups on the sides are commonly formed, are readily converted to COCl groups, and can react with a variety of reagents possessing the appropriate functional group. Also by virtue of the double bonds in the hexagonal rings, the tubes condense with nitrenes (to form aziridino compounds with UV light), react with carbenes (to form cyclopropane derivatives), are reduced (by Birch reduction), form metal derivatives which can be reacted further, react with aryl diazonium compounds (arylation), undergo 1,3-dipolar cycloaddition reactions (e.g. to form pyrrolidine derivatives), undergo nucleophilic reactions, radical reactions, and halogenation. J.L. Musgrave and coworkers [Khabashesku et al. *Acc Chem Res* **35** 1087 2002] have prepared ‘fluoronanotubes’ by direct fluorination of SWCNTs, and the fluoronanotubes dispersed in THF reacted with alkyl-Li reagents, Grignard reagents, alkoxides, hydrazine and  $\omega$ -diamines to form a variety of useful functionalised tubes. Sun and coworkers [Fu et al. *Nano Lett* **1** 439 2001] functionalised SWCNTs with lipophilic and hydrophilic dendra and showed that they can be defunctionalised in homogeneous solutions by base- and acid- catalysed hydrolysis demonstrating the existence of ester linkages. Absorption in the UV/VIS spectra of these functionalised tubes obey Beer’s law demonstrating no aggregation effects, and SEM analysis was not successful due to the extent of substitution, whereas the de-functionalised SWCNTs and MWCNTs gave successful SEM imaging. SWCNTs attached to a silicon substrate can be used as very small highly sensitive chemical sensors or memristors for gases. They also adsorb chemicals such as alcohols, aromatics, amines and phosphonates, some of which do not readily desorb and slow the process down unless they are removed by other means such as high temperatures which tend to be time consuming and degrade the sensor. Masel and coworkers [Salehi-Khojin et al. *Science* **329** 1327 2010, reviewed by H. Birch, Electric shock resets nanotube sensor, *RSC Chemistry World* **7** (10) 2010] showed that current induced voltage above the Poole-Frenkel conduction threshold (>12V) desorbed most molecules effectively after 3 hours and the sensors returned rapidly to baseline.

**SWCNT—(CO<sub>2</sub>H)<sub>n</sub>, 89-90% SWCNT (carbonaceous purity), bundle dimension: diam 4-5nm x length 0.5-1.5 $\mu\text{m}$ .** The SWCNTs were produced by EA discharge [Journet et al. *Nature* **388** 756 1996, Itkis et al. *J Phys Chem B* **108** 12770 2004] or the HiPco procedure (see above). SWCNT-CO<sub>2</sub>H is prepared as follows: HiPco SWCNTs (0.2g) are sonicated in a 3:1 mixture of concentrated H<sub>2</sub>SO<sub>4</sub>/concentrated HNO<sub>3</sub> (80ml) for 4 hours, diluted with deionised H<sub>2</sub>O and filtered through a 0.2 $\mu\text{m}$  pore acid resistant membrane. The solid is dried at ~25°, then sonicated in a solution of 4:1 concentrated H<sub>2</sub>SO<sub>4</sub>/30% H<sub>2</sub>O<sub>2</sub> for 15 minutes, diluted with excess deionised H<sub>2</sub>O, filtered again and dried *in vacuo* at 25° to give 79w/w% of ‘shortened’ product (85% have length < 600nm, i.e. ~0.6 $\mu\text{m}$ ). [Zhao et al. *Adv Funct Mater* **14** 71 2004.] Solution phase NIR spectroscopy against a standard [see Itkis et al. *Nano Lett* **3** 309 2003], gave a carbonaceous purity of 80-90% and 3-6 atom% carboxylic acid. Its solubility is 1mg/ml in DMF, and 0.1mg/ml in H<sub>2</sub>O. Metal content is 5-10%. This is suitable for further reactions.

**SWCNT—(COCl)<sub>n</sub>.** The preceding shortened SWCNT-acid (100mg) [see also Liu et al. *Science* **280** 1253 1998] is stirred in SOCl<sub>2</sub> (20ml) containing DMF (1ml) at 70° for 2 hours, centrifuged, the brown coloured supernatant is decanted, the sediment is washed with anhydrous THF, centrifuged again; the pale yellow supernatant is decanted and the sedimented SWCNT-(COCl)<sub>n</sub> is dried at ~25° *in vacuo*. This product is reactive and should be used immediately. [Chen et al. *Science* **282** 95 1998]. Alternatively, the preceding acid (12mg) in dry DMF (20ml) is sonicated for 30 minutes to give a homogeneous suspension, to which is added redistilled oxalyl chloride (0.4ml) dropwise at 0° under N<sub>2</sub>. The mixture is stirred at 0° for 2 hours, then at 70° and stirred overnight to remove excess of oxalyl chloride, and dried *in vacuo*. This product is reactive and should be used immediately. [Itkis et al. *Nano Lett* **3** 309 2003.]

**SWCNT—(CONH<sub>2</sub>)<sub>n</sub>, 89-90% SWCNT (carbonaceous purity), bundle dimension: diam 4-6nm x length 0.7-1.0µm, average diameter of individual SWCNT being 1.4nm ±0.1nm.** The preceding cooled SWCNT-(COCl)<sub>n</sub> in dry DMF (~0°) is treated with liquid NH<sub>3</sub> (~0.5ml) and stirred while the temperature rose to ~25°, and kept there with stirring for 24 hours. Dry N<sub>2</sub> is bubbled through the mixture to remove excess of NH<sub>3</sub>, excess of dry THF is added and the mixture is centrifuged, the supernatant is decanted off, the sediment is washed with dry THF, centrifuged again and the SWCNT-NH<sub>2</sub> residue is dried *in vacuo* at ~25°. The extent of labeling is the same as the original acid, i.e. 3-6 atom % (amide groups) and metals ~6-8%. Their solubilities are 0.5 to 1.0mg/ml, in each of the alcohols, acetone and DMF; and can be functionalised with any reagent that will react with amide groups.

**SWCNT-(ODA)<sub>n</sub> {SWCNT-[CONHCH<sub>2</sub>(CH<sub>2</sub>)<sub>16</sub>CH<sub>3</sub>]<sub>n</sub>, octadecylamide functionalised}, 89-90% SWCNT (carbonaceous purity), bundle dimension: diam 2-10nm x length 0.5-2.0µm, average diameter of individual SWCNT being 1.4nm ±0.1nm.** It is prepared from the above SWCNT-(COCl)<sub>n</sub> (obtained from 100mg of the acid) by mixing with octadecylamine (ODA, 2g, m 55-57°) and heating at 90-100° for 96 hours under dry N<sub>2</sub>, and cooled. Excess of ODA is removed by washing with EtOH four times (5-10 minutes sonication at 40 KHz) by centrifugation/decantation, the sediment is dissolved in CH<sub>2</sub>Cl<sub>2</sub>, filtered, the black coloured filtrate is evaporated to dryness in a rotavap and the residue is dried *in vacuo* at ~25° to give >60% yield based on shortened SWCNTs. Note that the reaction of SWCNTs-(COCl)<sub>n</sub> and excess of ODA in toluene at ~25° for several days gives only traces of product and the success of the former procedure was attributed to expansion and defoliation of the SWCNTs bundles to give the more reactive individual nanotubes [Chen et al. *Science* **282** 95 1998]. The extent of labeling in the commercial sample is 30-40wt% (ODA) and contains 4% of metals. The amide is soluble in CHCl<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub>, \*C<sub>6</sub>H<sub>6</sub>, toluene, and the solubility in CS<sub>2</sub> or THF is >1mg/ml.

The solubility of the ODA derivative made it amenable to purification, and R.C. Haddon and coworkers purified it by gel permeation chromatography (SEC) using Styragel HMW7 [Nyogi et al. *J Am Chem Soc* **123** 733 2001] and PLgel MIXED-A [Zhao et al. *J Am Chem Soc* **123** 11673 2001], both being polystyrene divinylbenzene resins. The latter proved to be a superior gel (300 x 7.5mm column), and using THF as eluent three bands were separated at a flow rate of 0.5ml/minute. The first band contained 74% of SWCNTs-ODA (as detected by AFM, UV and NIR) [MW range 2000 to 4 x 10<sup>7</sup>, particle size 15-20µm, retention time 8 minutes], the second band with retention time of 9 minutes contains mostly nanoparticles with traces of SWCNTs, and the third band with retention time of 19 minutes contained amorphous carbon. IR, Raman and UV spectra confirmed the ODA component, and <sup>1</sup>H NMR (200MHz, CDCl<sub>3</sub>) demonstrated bands characteristic of CH<sub>2</sub> groups and the terminal CH<sub>3</sub> group. The nature (e.g. such as broadness) of the bands was indicative that both ionic (charge transfer) and covalent interactions were occurring.

**SWCNT-{CONH-*p*-[C<sub>6</sub>H<sub>3</sub>(*m*-SO<sub>3</sub>H)-*p*-NHC<sub>6</sub>H<sub>3</sub>(*m*-SO<sub>3</sub>H)]<sub>m</sub> }<sub>n</sub> [SWCNT-*n*-poly-*p*-aminobenzene-*m*-sulfonic acid, SWCNT-(CO-PABS)<sub>n</sub>], 75-85% SWCNT (carbonaceous purity), average diameter of individual SWCNT being 1.4nm ±0.1nm, average PABS Mw ~400-600g/mol.** The PABS polymer was prepared by mixing *m*-aminobenzenesulfonic acid (ABS, 0.865g) and aniline (15-20mol% of ABS, as inhibitor of polymerisation) and 1M HCl with ammonium persulfate as oxidant, stirred at 0° for 6 hours, concentrated at ~25° (*in vacuo*), filtered and the solid was washed with Me<sub>2</sub>CO. This was dissolved in H<sub>2</sub>O, and the aqueous solution was slowly added to a large excess of Me<sub>2</sub>CO. The black solid was filtered off and dried at ~25° (*in vacuo*) to give PABS (340mg, 40%, Av M<sub>w</sub> ~400-600 g/mol), which was identified by its UV spectrum that has λ<sub>max</sub> at 290 and 510nm in aqueous 1N NaOH [Roy et al. *Synth Met* **100** 233 1999]. SWCNT-(CO-PABS)<sub>n</sub> was

obtained from SWCNT-(COCl)<sub>n</sub> [prepared as above by the oxaly] method from HiPco SWCNTs (12mg, Carbon Nanotechnologies Inc.) in dry DMF (~20ml) by mixing with PABS (120mg) in DMF (~50ml) and stirred at 100° for 5 days. After cooling to 25° the solid was filtered through a 0.2mm pore-size membrane and washed thoroughly with DMF and EtOH, and dried *in vacuo*. The black SWCNT-(CO-PABS)<sub>n</sub> (57mg) on the membrane was collected and dried *in vacuo* overnight. Its solubilities are 0.05mg/ml in EtOH, 0.1mg/ml in DMF and 5.0mg/ml in H<sub>2</sub>O. The commercial product has 65% (PABS) and 4% metals. The water soluble graft polymer had a much higher conductivity (5.6 x 10<sup>-3</sup> Scm<sup>-1</sup>) than PABS (5.4 x 10<sup>-7</sup> Scm<sup>-1</sup>), and IR spectrum consistent with an amide bond; the <sup>1</sup>H NMR (300MHz, D<sub>2</sub>O) exhibited a very weak broad spectrum compared with the sharp signals of PABS itself, characteristic of the effect of the ring currents in the nanotubes. The UV/VIS/MIR spectrum showed the presence of the interband transitions of the semiconducting SWCNTs and an absorption at 17,750 cm<sup>-1</sup> due to the PABS moiety. [Zhao et al. *Adv Funct Mater* **14** 71 2004.] Solution phase NIR spectroscopy against a standard, provided a carbonaceous purity of 80-90% — a procedure which has an accuracy of ~3% [see Itkis et al. *Nano Lett* **3** 309 2003].

**SWCNT-(PEG)<sub>n</sub> {SWCNT-[COO(CH<sub>2</sub>CH<sub>2</sub>-O)-<sub>m</sub>-H]<sub>n</sub>, polyethylene glycol functionalised}, 80-90% SWCNT (carbonaceous purity), bundle dimension: diam 4-5nm x length 0.5-0.6µm, average diameter of individual SWCNT being 1.4nm ±0.1nm, PEG Mw ~600g/mol.** The above SWCNT-(COCl)<sub>n</sub> prepared from 30mg of SWCNT-CO<sub>2</sub>H is mixed with PEG (250mg, m 20-25°) and heated under N<sub>2</sub> at 75° with vigorous stirring for 48 hours. The mixture is cooled to 25° extracted with CHCl<sub>3</sub> several times, filtered, and the dark coloured solution is repeatedly precipitated with EtOH to give SWCNT-(PEG)<sub>n</sub> which is collected (filtration or centrifugation) and dried *in vacuo*. Its solubility in H<sub>2</sub>O is high (5.0mg/ml) and has ~30wt% (PEG) and ~6% trace metals. The ester function can be identified by IR (ν<sub>max</sub> ~1700 cm<sup>-1</sup>), but the <sup>1</sup>H NMR signals are weak and broad (see above). The ester-free SWCNTs can be recovered by acid- and base- catalysed hydrolysis [see above Fu et al. *Nano Lett* **1** 439 2001]. Solution phase NIR spectroscopy against a standard, provides a carbonaceous purity of 80-90% — a procedure which has an accuracy of ~3% [see Itkis et al. *Nano Lett* **3** 309 2003].

**SWCNT-(*p*-C<sub>6</sub>H<sub>4</sub>-R; R = F, Cl, Br, I, SO<sub>3</sub>H, CO<sub>2</sub>H, NO<sub>2</sub>, *n*-butyl, *t*-Bu or CO<sub>2</sub>Me)<sub>n</sub>.** These were prepared by ‘on water’ functionalisation of bundled SWCNTs. The term ‘on water’ refers to the water-based reactions of water-insoluble organic substrates [see K.B. Sharpless and coworkers in Narayan et al. *Angew Chem, Int Ed* **44** 3275 2005 and Klijn & Engberts *Nature* **435** 746 2005]. This technically involves the reaction of *p*-substituted benzene diazonium compounds with the hexagonal rings on the walls of SWCNTs in aqueous medium and represents a “green” or “environmentally friendly” process. The following are optimal conditions for functionalisation. HiPco SWCNTs (10mg) and deionised H<sub>2</sub>O (30ml) were homogenised in a flask (100ml) by stirring at medium setting for 30 minutes, then heating at 80° with the substituted aniline (4 equivalents/SWCNT) and isoamyl nitrite (2 equivalents/SWCNT) with vigorous stirring (stirrer bar) under a reflux condenser overnight. The mixture is then cooled, filtered through a 0.45µm Teflon filter, the filter cake is washed with deionised H<sub>2</sub>O, and Me<sub>2</sub>CO until the filtrate is clear. The cake is collected, sonicated in DMF (25ml) to remove any remaining organic compounds, collected by filtration (0.45µm Teflon filter) and rinsed with Me<sub>2</sub>CO to give the desired functionalised SWCNTs as evidenced by 20-30% weight increases. TGA, Raman [elevated D (diamondoid, ν<sub>max</sub> at 1290 cm<sup>-1</sup>)/G (graphic, ν<sub>max</sub> 1590 cm<sup>-1</sup>) band ratios] and UV-VIS-NIR spectroscopy, XPS, AFM and TEM confirmed the structures. It is interesting that TEM images (on a carbon grid) showed that the SWCNT bundles have smooth edges, whereas functionalised SWCNT bundles have ‘bumps’ all along the edges. These functionalised nanotubes should be useful for further functionalisation reactions [Price & Tour *J Am Chem Soc* **128** 12899 2006].

By using the above procedure SWCNT-(*p*-C<sub>6</sub>H<sub>4</sub>-CH<sub>2</sub>NH<sub>2</sub>) and MWCNT-(*p*-C<sub>6</sub>H<sub>4</sub>-CH<sub>2</sub>NH<sub>2</sub>) were prepared and successfully coupled, *via* their terminal amino group to the carboxy group of *N*(1)-carboxymethyl-thymine, to form thymine ends. These functionalised CNTs readily form stable *double* hydrogen bonds with other thymine groups of these CNTs. This induces controlled non-covalent self-assembled supramolecular aggregation of the CNTs in solvents that do not break hydrogen bonds such as CH<sub>2</sub>Cl<sub>2</sub>, and can form good dispersions in polar aprotic solvents such as DMF. All characterisations were performed using spectroscopic, analytical and microscopic techniques. [Quintana & Prato *Chem Commun* 6005 2009, reviewed by K. Davies *RSC Chemistry World* **6**(11) 6005 2009.] These properties can be of importance in *nanoelectronics*, or in

biological applications such as making patterned active substrates for neuronal growth [Cellot et al. *Nat Nanotechnol* **4** 126 2009].

**I-SWCNT**,  $\mathbf{I} = \{-\text{COOCH}_2\text{-[3,5-di(hexadecyloxy)phenyl]}\}_n$  and **I<sub>PEG</sub>-SWCNT**,  $\mathbf{I}_{\text{PEG}} = \{-\text{COOCH}_2\text{-[3,5-di(methyltriglycoloxy)phenyl]}\}_n$ . These are respectively hydrophobic and hydrophilic SWCNTs which were prepared as described by Y.-P. Sun and coworkers [Fu et al. *Nano Lett* **1** 439 2001] by stirring vigorously SWCNT-(COCl)<sub>n</sub> (30mg, see above) and carefully dried **I** (250mg) under N<sub>2</sub> at 75°/48 hours. The cooled mixture is extracted several times with CHCl<sub>3</sub>, the combined dark-coloured extracts were repeatedly precipitated with EtOH to give **I-SWCNT** with <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) which had δ at 0.88 (t, *J* = 6.5Hz) and typically weak broad bands at 1.1-1.5, 1.6-1.8, 3.5-3.9, 4.0-4.2 and 6.0-6.5 caused by the effect of the CNT. **I<sub>PEG</sub>-SWCNT** was prepared in a similar manner except that its solubility in H<sub>2</sub>O allowed further purification by dialysis for several days against deionised H<sub>2</sub>O (dialysis tubing with Mr ~100,000 cut off to remove PEG). Dendron **I** was prepared by reaction of methyl 3,5-dihydroxybenzoate and hexabromodecane followed by reduction with LAH, and Dendron **I<sub>PEG</sub>** was obtained from the same benzoate and triethylene glycol monoethyl ether in the presence of Ph<sub>3</sub>P and diethyl azidodicarboxylate in THF, followed LAH reduction. The corresponding hydrophobic **I-MWCNT** and hydrophilic **I<sub>PEG</sub>-MWCNT** dendra were similarly prepared from reacting SWCNT-(COCl)<sub>n</sub> with the benzylic OH group in the dendra, and all were characterised by <sup>1</sup>H NMR, UV-Vis spectroscopy, TGA and SEM as well by defunctionalisation of homogeneous solutions under acid- and base- catalysed reaction conditions.

**MWCNT-(PPEI-EI)<sub>n</sub>**. These functionalised MWCNTs were prepared in two ways. *Firstly*, the MWCNTs (20mg) were refluxed with SOCl<sub>2</sub> (5ml) for 24 hours and evaporated *in vacuo*, then the co-polymer PPEI-EI [200mg, poly(propionylethylenimine-co-ethylenimine) [as prepared by Y.-P. Sun and coworkers *Macromolecules* **32** 8747 1999, *Photochem Photobiol* **66** 301 1997, *Chem Commun* 2699 1996]] was added and heated at 160-180° for 12 hours under N<sub>2</sub>, cooled, repeatedly extracted with CHCl<sub>3</sub> and the dark combined extracts were precipitated with hexane. The isolated solid was dissolved in deionised H<sub>2</sub>O and dialysed against fresh H<sub>2</sub>O (dialysis tubing with Mr ~100,000 cut off to remove PPEI-EI) for 3 days. Further purification was by re-precipitation from CHCl<sub>3</sub> solution into hexane, and drying the dark **MWCNT-(PPEI-EI)<sub>n</sub>** solid which had <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) with broad weak bands at δ 0.8-1.4, 1.9-2.6, 2.7-2.9, 3.1-4.3. *Secondly*, (without forming the COCl derivative) by directly heating the MWCNTs (20mg) with PPEI-EI (200mg) at 160-180° for 12 hours under N<sub>2</sub> and worked up as above provided **MWCNT-(PPEI-EI)<sub>n</sub>** which had <sup>1</sup>H NMR (500MHz, CDCl<sub>3</sub>) with broad weak bands at δ 0.8-1.4, 1.9-2.6, 2.7-2.9, 3.1-4.3. Raman, UV-VIS spectroscopy, SEM, TEM and TGA showed that both procedures were effective in producing functionalised MWCNTs which were readily soluble in common organic solvents and in H<sub>2</sub>O. The latter method may have caused amidation as well as ionic interaction between the amino-polymer and the CNT. [Lin et al. *J Phys Chem B* **106** 1294 2002.]

**SWCNT[=CH]<sub>2</sub>N-CH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>NH<sub>3</sub><sup>+</sup>. HCl]<sub>n</sub> and **MWCNT[=CH]<sub>2</sub>N-CH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>-OCH<sub>2</sub>CH<sub>2</sub>NH<sub>3</sub><sup>+</sup>. HCl]<sub>n</sub>**. These salts are very soluble in H<sub>2</sub>O and solutions are stable for more than a month at concentrations of 20mg and 12mg per ml respectively. The free amino terminal group is a useful handle for attachment to physiologically active molecules (see below). They are formed in a 1,3-dipolar cycloaddition reaction between the C-C bonds of the CNTs with the N-glycine N atom and formaldehyde. The glycine in this case is BocNHCH<sub>2</sub>CH<sub>2</sub>O-CH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>NHCH<sub>2</sub>CO<sub>2</sub>H and is prepared as follows: BocNHCH<sub>2</sub>CH<sub>2</sub>O-CH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>NH<sub>2</sub> (30mmol) in dioxane (20ml) at 0° is treated dropwise with a solution of benzyl bromoacetate (2.3g, 10mmol) in dioxane (30ml) during 1 hour, and the mixture is stirred overnight. The solvent is evaporated off *in vacuo*, the residue is dissolved in H<sub>2</sub>O (70ml) and extracted with EtOAc (3 x 50ml), the combined organic phase is dried (Na<sub>2</sub>SO<sub>4</sub>), evaporated *in vacuo*, and the residue is purified by chromatography on Silica gel [NM Kieselgel 60 (70-230 mesh)] and eluted with 1:1 EtOAc/petroleum ether then pure EtOAc to give *N-Boc-aminoethoxyethoxyethylaminoacetic* as an oil. To a solution of this oil (5.05mmol) in MeOH (50ml) is added 10% Pd/C (50mg), the mixture is stirred under H<sub>2</sub> for 24 hours, the catalyst is filtered off (through Celite), the solvent is evaporated and the residue is triturated with dry Et<sub>2</sub>O to give *aminoethoxyethoxyethylaminoacetic acid* as a pure white solid (1.6g, 99%, 5.05mmol), **m 105-106°** with the expected elemental (C, H and N) analyses. The acid has IR-DRIFT (KBr) with  $\nu_{\text{max}}$  at 3250, 2970, 1706, 1620, 1540, 1365, 1115, 686, 590, 480 cm<sup>-1</sup>; the <sup>1</sup>H NMR (200MHz, CDCl<sub>3</sub>, TMS) has δ at 1.40 (s, 9H), 3.22 (m, 2H),**

3.49 (t,  $J = 5.1$  Hz, 2H), 3.64-3.53 (m, 8H), 3.79 (bt, 2H), 5.54 (bt, 1H), 6.23 (bs, 1H), 8.21 (bs, 1H); the  $^{13}\text{C}$  NMR (50MHz,  $\text{CDCl}_3$ , TMS) has  $\delta$  at 170.5, 156.2, 79.1, 70.4, 70.3, 70.1, 66.6, 49.8, 46.8, 40.4, 28.6; EI-MS found  $m/z$  306 ( $\text{M}^+$ ). [Kordatos et al. *J Org Chem* **66** 4915 2001.]

For **functionalisation**, a suspension of full length SWCNTs or MWCNTs [diameter 20-30nm, from Carbon Nanotechnologies, Inc USA ([www.cnanotech.com](http://www.cnanotech.com)), and Nanostructured & Amorphous Materials Inc USA, ([www.nanoamor.com](http://www.nanoamor.com)) respectively] in DMF is treated with the preceding acetic acid and paraformaldehyde, and the mixture is heated at  $130^\circ$  for 96 hours. Unreacted material is removed by filtration, the filtrate is evaporated, the residue is dissolved in  $\text{CHCl}_3$ , washed with  $\text{H}_2\text{O}$ , dried, evaporated, redissolved in  $\text{CHCl}_3$ , precipitated with  $\text{Et}_2\text{O}$ , collected (on a 0.45mm Teflon filter) and washed several times with  $\text{Et}_2\text{O}$  to give the functionalised CNTs in  $\sim 10\%$  yields based on the amount of starting CNTs. They are soluble in solvents such as  $\text{CH}_2\text{Cl}_2$ ,  $\text{CHCl}_3$ , toluene and  $\text{Me}_2\text{CO}$ . Removal of the *N*-Boc group is achieved by dissolving the previous functionalised SWCNTs or MWCNTs in  $\text{CH}_2\text{Cl}_2$ , dry HCl gas is bubbled through the solution whereby the **CNT-chloride hydrochloride salts** (desired materials) separate out. They are collected (or the solvent is evaporated), dissolved in MeOH and precipitated with dry  $\text{Et}_2\text{O}$ . The desired products, as identified by TEM showed that the functionalised SWCNTs and MWCNTs have 10-50nm and 20-30nm mean diameters respectively, their  $^1\text{H}$  NMR spectra are similar with the signals from the oligoethylene glycol chains appearing as broad peaks at  $\sim 3.6$  ppm, and absence of the Boc methyl groups which would have been at 1.2 ppm from TMS. [Georgakilas et al. *Chem Commun* 3050 2002.] These functionalised CNTs were successfully coupled *via* their terminal amino group to the carboxy group of *N*(1)-carboxymethyl-thymine to produce thymine ends. The thymine functionalised CNTs readily form stable double hydrogen bonds with other thymine groups of these CNTs. This induces controlled non-covalent self-assembled supramolecular aggregation on the CNTs in solvents that do not break hydrogen bonds such as  $\text{CH}_2\text{Cl}_2$ , and can form good dispersions in polar aprotic solvents such as DMF. All characterisations were performed using spectroscopic, analytical and microscopic techniques. [Quintana & Prato *Chem Commun* 6005 2009, reviewed by K. Davies *RSC Chemistry World* **6**(11) 6005 2009.] These properties can be of importance in *nanoelectronics* or in biological applications such as making patterned active substrates for neuronal growth [Cellot et al. *Nat Nanotechnol* **4** 126 2009.]

SWCNT[ $=(\text{CHR}')(\text{CH}_2\text{N-R}'')$ ], where  $\text{R}' = \text{H}$ , 4-MeOC<sub>6</sub>H<sub>4</sub>- or 2-pyrenyl-, and  $\text{R}'' = -(\text{CH}_2\text{CH}_2\text{O})_3\text{CH}_3$  or  $-(\text{CH}_2)_7\text{CH}_3$ ]<sub>*n*</sub> and MWCNT[ $=(\text{CHR}')(\text{CH}_2\text{N-R}'')$ ], where  $\text{R}' = \text{H}$ , 4-MeOC<sub>6</sub>H<sub>4</sub>- or 4-pyrenyl-, and  $\text{R}'' = -(\text{CH}_2\text{CH}_2\text{O})_3\text{CH}_3$  or  $-(\text{CH}_2)_7\text{CH}_3$ ]<sub>*n*</sub> were prepared by 1,3-dipolar cycloaddition reactions with short oxidised SWCNTs or purified MWCNTs as in the preceding entry and using paraformaldehyde, 4-methoxybenzaldehyde or 4-formylpyrene to form the substituted pyrrolidines with the double bond(s) of the CNTs. They were characterised by UV-VIS, Raman and  $^1\text{H}$  NMR spectroscopy and by TEM. The fluorescence of the pyrene derivatives were of interest, and all provide means for the preparation of nanocomposites. [Georgakilas et al. *J Am Chem Soc* **124** 760 2002, Calvert *Nature* **399** 210 1999.]

## GRAPHENE MATERIALS

**Graphene** is the name given to a flat monolayer of carbon atoms packed in a two-dimensional (2D) honeycombe lattice of six-membered carbon rings. It is the unit block (2D sheet) of graphitic materials which stack into multilayers to form graphite (3D), rolls to form nanotubes (1D) or wraps to form buckyballs (0D)[see above]. ‘Graphenes’ are defined as single-, double-, and few (3 to <10) carbon layers (2D), with characteristic electrical properties not found in graphite. *Graphene* is the accepted name for a single, or at most a double carbon thick layer which is  $\approx 5\text{\AA}$  thick. Single layer 2D crystals have been prepared (see below) that have high crystal quality where charge carriers can travel thousands of atomic distances without scattering. Its electrical behaviour includes a pronounced ambipolar electric field effect in which charged carriers could be tuned continuously between electrons and holes at concentrations  $n$  as high as  $10^{13}\text{ cm}^{-2}$ , with mobilities  $\mu$  greater than  $15,000\text{ cm}^2\text{ V}^{-1}\text{ s}^{-1}$  under laboratory conditions. Films of 5 layers would be considered as ‘bulk material’, and films with more than 100 layers are considered as thin films of 3D graphite (see graphite [7782-42-5] below). [Geim & Novosolov *The Rise of Graphene: Nat Mat* **6** 183 2007.]

Graphene films were obtained by mechanical exfoliation (repeated peeling) of highly oriented pyrolytic graphite. The 2D crystallites were collected from a fresh surface of a layered crystal using few-layered graphene, FLG, by rubbing against another surface (any solid is suitable, but an oxidised silicon wafer, i.e.  $\text{SiO}_2$  would be also useful later for identification). This allows flakes of graphene to be attached to the surface for characterisation. This approach was very reproducible, has also been used for preparing thicker films ( $\geq 3\text{nm}$ ) which were up to  $100\mu\text{m}$  across, as well as FLGs of up to  $10\mu\text{m}$  in size and visible to the naked eye.

Preliminary identification of single layer films among the resulting flakes was done with an optical microscope. These became visible when they were on top of an oxidised silicon wafer (see above) because even a monomolecular layer adds enough to the optical path of reflected light that the interference colour changes reveal the presence and contours of the film. The process is simple and the 2D crystallites can be identified in about half an hour. Further identification of selected single layers is then done by AFM. It was pointed out that only a small number of single films among the many thicker films were produced by this method, they could not be identified by TEM, but they were transparent to visible light on surfaces other than SiO<sub>2</sub> wafers, and thus require AFM and high resolution AFM imaging for definitive identification and are difficult to find by scanning surfaces at random. These atomically thin sheets, particularly very large 2D molecules, unprotected from the environment are stable in air under laboratory conditions, have high crystal quality and maintain macroscopic continuity. [Novoselov, Geim and coworkers, Electric field effects in atomically thin carbon films, *Science* **306** 666 2004, Two dimensional atomic crystals, *Proc Natl Acad Sci USA* **102** 10451 2005.]

The sp<sup>2</sup> character of all the C-C bonds of graphene confer remarkable electric conductivity to it due to the delocalisation of electrons, and it is difficult to stop the flowing current at will. Geim and Novoselov and their coworkers [Elias et al. Control of Graphene's properties by reversible hydrogenation: Evidence for Graphane, *Science* **323** 610 2009] have succeeded in hydrogenating the double bonds of graphene to make **graphane** which now has C-C bonds with sp<sup>3</sup> character. The reduction transforms graphene from the highly conductive zero-overlap semimetal into an insulator. Reduction was achieved with a hydrogen-helium mixture (10% H<sub>2</sub>) at low pressure (0.075mm Hg) and the direct current plasma ignited between two aluminium electrodes with the samples 30cm away from the discharge zone to avoid damage by energetic ions. Graphene and graphane have distinctly different Raman properties. The reaction is reversed by annealing graphane at 450° (higher temperatures damage graphene) in an argon atmosphere for 24 hours. Thus the original metallic state, lattice spacing, Raman and TEM properties, and even the quantum 'Hall effect' are restored. Note that when the thin film is on a substrate, reduction takes place on the upper face of the film, and for more effective reduction they used free-standing graphene membranes. If one can imagine a magic pencil which can remove hydrogen from graphane and drawing channels of graphene into it, better control of current flow can be achieved [see review by J. Urquart *RSC Chemistry World* **6** (3) 23 2009].

**Cutting of single-layer graphene** (SLG) by anisotropic etching using thermally activated Ni nanoparticles has been developed by Jarillo-Herrero and coworkers [Campos et al. *Nano Lett* **9** 2600 2009, reviewed by L. Brindley, Cutting Graphene to ribbons, *RSC Chemistry World* **9** (8) 2009]. Unlike previous methods for making graphene nanoribbons and the like, their procedure makes SLG nanoribbons and other SLG structures (e.g. triangles, rectangles) in which all cuts (sub-10nm width) are oriented with the same edge-chirality, i.e. aligned along a single crystallographic directions with no hanging edges. Thus the cuts are straight, and always at an angle of 60° and/or 120°. With graphite and FLGs, etching occurs with chirality changing angles of 30°, 90°, and 150°, as well as 60° and 120°. Their method is described here in some detail as it involves **purification procedures useful for SLG preparation**. To ensure reliable cuts in nanoparticle assisted etching, the SiO<sub>2</sub>/Si substrates (wafers, see above) have to be cleaned with Me<sub>2</sub>CO and *iso*-PrOH, and submitted to UV illumination for 5 minutes to remove organic material and for making the SiO<sub>2</sub>/Si surface more hydrophilic to ensure proper wetting during the spin-coating deposition of Ni from solution. Graphene is exfoliated on this clean surface using semiconductor grade tape. The SLG on the clean substrate is identified by optical spectroscopy (see above), AFM (typically SLGs are at a height of 0.8nm) and Raman spectroscopy (graphene's G' 2D peak, see above). The tape residue is removed by heating the SLGs in a quartz tube at 500° for 15 minutes under an argon:H<sub>2</sub> flow (850:150 sccm). After this heat 'cleaning', an aqueous solution of NiCl<sub>2</sub> (2.4mg/ml) is centrifuged onto this surface at 1800rpm for 60 seconds, then heated at 90° for 10 minutes on a hot plate to eliminate H<sub>2</sub>O. The NiCl<sub>2</sub> treated sample in an argon:H<sub>2</sub> flow (850:150 ratio) is firstly annealed at 500° for 20 minutes to convert the salt into Ni nanoparticles, and then the etching is carried out at 1000° for 25 minutes in the same gas flow. The best etching conditions require a slow heating/cooling rate of ~50°/minute. During the etching the Ni nanoparticle (with H<sub>2</sub> adsorbed) furrows through the SLG by catalytically hydrogenating the carbon atoms in its path and converting each atom into a molecule of methane. With SLGs, but not with FLGs or graphite, the cutting paths never cross each other; as soon as the Ni nanoparticle comes within a 10nm distance of a preformed path it changes direction at a 60° or 120° angle. The purification procedure and the Ni concentration stated are very critical if the formation of CNTs are to be avoided. This opens many future directions for studying SLG nanoribbons and other shapes, constrictions and quantum dots with crystallographic edges. H-H. Ahn and B.H Hong and coworkers [Bae, et al, *Nature Nanotechnology* **5** 574 2010] have pressed rectangular graphene films (76cm diagonal) with a roller against an adhesive polymer support and the copper

etched away. The graphene is pressed with rollers onto a polyethylene terephthalate substrate and the polymer adhesive is released on heating. In this way further layers of graphene can be added. When treated with  $\text{HNO}_3$ , the graphene sheet acts as a transparent electrode for touchscreen devices. It is tougher and has better transparency than indium-tin oxide (ITO) electrodes [see below, and review by S. Hadlington RSC *Chemistry World* 7 (8) 22 2010].

**Stacked Graphene Platelet Nanochips (stacked graphene nanofilms, SGNF heat treated) and Platelet Nanofibres (SGNF acid washed)** are commercially available as black powders with 40-50nm mean width, 0.1-10 $\mu$ m length, 0.3g/cm<sup>3</sup> density, and 120m<sup>2</sup> of surface area. The electrical resistivities are 55 $\mu$ Wcm and 120 $\mu$ Wcm for the chips and fibres respectively. These SGNFs are used in nanotechnology research. [US Patents 6,995,115 and 7,001,586.]

**Graphite** [7782-42-5] **m 3652-3697°**, **d<sup>25</sup> 2.09, 2.23**. It is available commercially in various forms such as *flakes* (mesh 75+), *rods* (6mm x 150mm, low density and high density, 3mm x 150mm low density), *powder* ( $\leq 20\mu\text{m}$ ,  $\leq 45\mu\text{m}$ , and  $150\leq\mu\text{m}$ ), *nanofibres* (I.D. 0.5-10nm x O.D 80-200nm x L. 0.5-20 $\mu\text{m}$ , bulk density 0.06-0.08g/ml), *graphitised nanopowder* (particle size <200nm with large mesopores, 10% graphite lattice content, surface area 70m<sup>2</sup>/g, average pore  $\sim 137\text{\AA}$  and d 1.818g/cm<sup>3</sup>) and *platelet nanofibres* (prepared by catalytic CVD then demineralised by HCl, with surface area 80m<sup>2</sup>/g, particle width  $\sim 100\text{nm}$  and particle length  $\sim 2.5\mu\text{m}$ , TEM). [For purification by acid treatment see entry [7782-42-5] in Chapter 5, Inorganic Compounds.] The applications of graphite include the manufacture of 'lead' pencils, lubricants, polishing, pigments, explosives, commutator brushes, carbon anodes, and arc-lamp carbon among many others. **Avoid breathing it** as it can cause coughing, decreased lung function and fibrosis, **use a mouth mask**. [See Holliday et al. in *Comprehensive Inorganic Chemistry* vol 1, Bailar Jr et al. Eds, Pergamon Press, Oxford, pp1250-1294 1973; ISBN 9780080172750.]

### SELF ASSEMBLED MONOLAYERS (SAMs).

SAMs have been of considerable interest for some time [see Chechik, Crooks and Stirling, Reactions and reactivity of self-assembled nanolayers, *Adv Mater* 12 1161 2001] and several useful ones have been prepared using a variety of materials and methods. Only a recent report will be described here, and it is of 2D carbon nanolayers prepared by Mo(IV) and Cu catalysed crosslinking of 1,4-bispropyne- (1), 1,4-bisacetylene- (2), and 1,4-bis(2,3,4-triacetylenephényl)propyne- (3) benzenes with each benzene ring further substituted with a 3-(triethoxysilyl)propylaminocarbonyloxymethyl— [(EtO)<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>-NH-COOCH<sub>2</sub>—] group which acts as a linker in the polymerisation process. These three organic compounds are connected by vacuum-driven Mo(IV) catalysed alkyne metathesis [cf Zhang & Moore *Adv Synth Catal* 349 93 2007], or oxidative Cu catalysed Hay-type coupling [cf Siemen et al. *Angew Chem Int Ed* 39 2633 2000] which polymerise them to form self-assembled monolayers on SiO<sub>2</sub>-, Si<sub>3</sub>N<sub>4</sub>- coated substrates (wafers) or quartz/glass slides when immersed in monomer solutions.

*Preparation and purification of substrates:* The freshly cleaned and dried SiO<sub>2</sub>-, Si<sub>3</sub>N<sub>4</sub>- coated substrates or quartz/glass slides are placed in a reaction tube containing a 15mM solution of (1) or (2), or (3) in toluene and Me<sub>3</sub>N (10mM) under N<sub>2</sub>, sealed and heated at 95-100° for 24 hours. The liquid is removed, the substrate is rinsed with toluene (1x), CH<sub>2</sub>Cl<sub>2</sub> (2x), and sonicated for 5 minutes in toluene. The rinse is repeated, the substrate is sonicated in MeOH for 5 minutes, the rinse is repeated again and the substrate is blown dried under a stream of N<sub>2</sub> ready for monolayer formation.

*Preparation and purification of polymerised coated substrate linked monolayers by Cu Coupling:* CuCl (20mg) and degassed Me<sub>2</sub>CO (3ml) in a reaction vial are treated with TMEDA (61 $\mu$ l), stirred under N<sub>2</sub> at 20° for 30 minutes and the preceding coated substrates are added. The mixture is purged with O<sub>2</sub> from a balloon and stirred under O<sub>2</sub> for 15 hours. The polymerised coated substrate is collected, rinsed with DMF (1x), with a solution of 0.1M sodium diethyldithiocarbamate in DMF (1x), with toluene (1x), with CH<sub>2</sub>Cl<sub>2</sub> (2x), sonicated in toluene for 5 minutes, the rinse is repeated, the substrate is finally sonicated for 5 minutes in MeOH, and the *polymerised linked substrate* blown dry with a stream of N<sub>2</sub>.

*Preparation and purification of polymerised coated substrate linked monolayers by alkyne metathesis:* A mixture of trisamidomolybdenum(IV) propylidyne (5.0mg) and *p*-nitrophenol (3.3mg) dissolved in trichlorobenzene (3.3ml, 1,2,4-) are placed in a vial, the coated substrate is added, the flask is sealed under a vacuum (5 Torr) for 22 hours. The substrate is collected. Rinsed with DMF (1x), 0.1M sodium diethyldithiocarbamate in DMF (1x), toluene (1x), CH<sub>2</sub>Cl<sub>2</sub> (2x), then sonicated in toluene for 5 minutes. The

rinse is repeated, the substrate is sonicated in MeOH for 5 minutes, the rinse repeated again and finally the *polymerised linked substrate* is blown dry in a stream of N<sub>2</sub>.

These monolayers (SAMs) are characterised by UV-Vis, fluorescence and Raman spectroscopy. This is made easy because of the presence of benzene cores within the acetylene and linker chains, these being made of poly(1,4-phenylene-1,3-butadiynylene) polymers. AFM then provides the topology, i.e. shapes (contours), thicknesses and aspect ratios (length and width), of the 2D SAMs. This is done at all stages including those described below.

*Lift-off and transfer of the SAMs:* The ‘lifting off’ is carried out by simple oxygen-reactive ion etching through a layer of ‘photoresist’ (e.g. 5µm x 5µm ‘aspect ratio’) patterned by photolithography. The ‘photoresist’ is placed on top of the monolayer which lies on top of the silicon substrates (support, i.e. the films of SiO<sub>2</sub>, or Si<sub>3</sub>N<sub>4</sub>, or glass). Selective etching from the support is done with concentrated HF which dissolves the silicon, leaving photoresist/SAM or photoresist/monolayer membrane hybrid which can be transferred onto other substrates and then freed from the ‘photoresist’ layer by sonication in Me<sub>2</sub>CO. Characterisation of the monolayers during these stages showed that very little, if any, changes in structure has occurred. The atomic force micrographs (AFMs) showed that the SAMs derived from (1) are monolayers with thicknesses of ≈ 1.5nm (average roughness of 0.18-0.27nm), and those from (2) and (3), unlike from (1), formed multilayers of 2.8nm and 4.7nm respectively (average roughness of 0.5nm and 0.3nm respectively). These SAMs, with aspect ratios being the same as the ‘photoresist’ can be transferred by contact onto a variety of surfaces such as films, pleated sheets, spheres, tubes, cones and any other shapes (followed by removal of the ‘photoresist’ by sonication in Me<sub>2</sub>CO). With a free-standing shape, a free standing silicon support can be used which is then removed by etching to give a free standing SAM. The monolayers are quite robust and can be stretched over ~440nm holes without tearing. This approach and these procedures that were developed by J.S. Moore, J.A. Rogers and coworkers [Schultz et al. *Poc Natl Acad Sci USA* **105** 7353 2008] have considerable merit for synthesising monolayer carbon networks of varying shapes which can be accessible for analysis and device applications.

## DIAMOND NANOMATERIALS

**Nanodiamond (ND)** [7782-40-3] **M 12.01** is commercially available in powder form with particle spherical size <10nm and ≥ 97% trace metal basis, bulk density 0.2-0.7g/ml with BET surface area of 200-450m<sup>2</sup>/g, and in powder form with particle spherical size <10nm and ≥ 95% trace metal, bulk density 0.17g/ml with BET surface area of 278-335m<sup>2</sup>/g. The diamond powder is made by *explosive detonation procedures* that are performed in stainless steel hermetic tanks (e.g. of 0.17m). Carbon condensed products [the HE (High Explosive) composite explosives] such as cast trotyl/cyclotrimethylene-trinitramine (TNT/RDX) in various proportions from 50/50 and increasing RDX, with various initial inert gas (N<sub>2</sub> or Ar) pressures were used to determine the best conditions to produce diamonds. Thirty minutes after the explosion, the pressure in the tank is slowly released, and the carbon product is washed out with water, allowed to settle (requiring quite a long period), the water is decanted off, and residual water is evaporated at 150°. This yielded solid carbonaceous soots which contain ultradispersed diamond (UDD, 2-15nm), together with amorphous carbon (4-25nm), graphite ribbons (<20nm), and spheres (2-4nm) as investigated by TEM, XRD, SAXS (small angle X-ray scattering) and Auger spectroscopy. Purification was carried out to remove amorphous and graphite-like carbon from the detonation soot by boiling with perchloric acid, and the maximum yield of diamond was identified by X-ray analysis. The best yield was ~7wt% obtained when TNT/RDX was 50/50 and initial pressure in the tank was ~7 atmospheres which optimised the yield of UDD. It was also shown that the UDD was thermally transformed to onion-like carbon (OLC) particles. [Kuznetsov et al. *Carbon* **32** 873 1994.] Annealing studies, in the temperature range of 750°—1900°, of detonation diamond (e.g. OLC) in an inert atmosphere (e.g. argon) affected the structure as shown by XRD and HRTEM as well as purified the material further. The IR showed that all the oxygenated groups (mainly OH, COOH and C=O) were removed, and that at the lower temperatures of ~700-800° a higher coverage of the diamond surface with π-bonds resulted. At higher temperatures the powder undergoes a phase transition from cubic diamond to graphite. [Chen et al. *Appl Phys Lett* **74** 3651 1999, Okotrub et al. *J Phys Chem A* **105** 9781 2001.] Previously, reactions of ND with free radicals lead to diamond alkylation [Nakamura et al. *Chem Commun* 900 2003]; and arylation with diazonium salts [Yeap et al. *Langmuir* **25** 185 2009, Liang et al. *ACS Nano* **3** 2288 2009] resulted in low yields of *functionalized nanodiamond*. However, with ND annealing at ~800°, the larger coverage of diamond surface with π-bonds allowed A. Kreuger and coworkers [Jarre et al. *Chem Commun* **47** 544 2011] to achieve stable covalent C-C bonding of aromatic groups using Diels-Alder reaction onto surface annealed ND. The key reaction involved the addition of *o*-quinodimethane [acting as diene, generated *in situ* from *o*-bis(bromomethyl)benzene and substituted

benzenes] to the  $\pi$ -bonds of ND in the presence of 18-crown-6 and KI in refluxing toluene for 72 hours. The substituent on the benzene ring can be *functionalised* further, e.g. an  $\text{SO}_3\text{H}$  group (formed by sulfonation of the benzene ring) can be converted into an SH group ( $\text{PPh}_3$ ,  $\text{I}_2$ ,  $^*\text{C}_6\text{H}_6$ , reflux, 24 hours) which can be made to react with *N*-dye substituted maleimide to form a condensed sulfide. The dye being fluorescent Oregon Green 488 which provided fluorescent diamond particles. The applications of such functionalisations are similar to those of functionalised fullerenes and CNTs including fluorescence labeling, drug delivery, magnetic sensing, electrochemistry, and composite materials among others.

### SOLVENT RESISTANT NANO FILTERS (SRNF)

During the past decade and a half considerable effort has been expended in developing filters for separating mixtures down to the molecular level. An alternative name to SRNF is organic solvent nanofiltration (OSN). This shows promise as an energy- and waste- efficient process. SRNF has been developed particularly for use with organic solvents but it also applies to aqueous media, with the *proviso* that the membranes may be different. The driving force is the pressure applied to the membranes which are classified as UF (ultrafiltration) with pressures of 1 to 5 bar, NF (nanofiltration) with pressures from 5 to 20 bar and RO (reverse osmosis) with pressures  $>10$  bar. In UF substances between 2nm and  $0.1\mu\text{m}$  are rejected, in NF particles or dissolved molecules smaller than 2nm are removed from solution, and in RO only solvents permeate.

SRNFs are essentially of two types, the organic polymer type or the ceramic type. The former are polymers which need to be cross-linked in order to increase their strength. A variety of polymers have been prepared and studied including the composites: PAN (polyacrylonitrile), PVDF [poly(vinylidene fluoride)], PI [polyimide (Matrimide) and (Lenzing P84)], PEI [poly(etherimide)], PA (polyamide), PAH [poly(amide hydrazide)], PSf (polysulfone), PES [poly(ethersulfone)], PEEK (polyether ether ketone), SPEEK (sulfonated PEEK), PPESK [poly(phthalazinone ether sulfone ketone)], SPPEEK (sulfonated PPESK), CA (cellulose acetate), and PBI (polybenzimidazole). The polymers are supported by an ultra-thin (submicron) barrier (TFC, thin film composite) on top of a chemically different polymer support. The TFCs are made of a variety of substances including PDMS (polydimethylsiloxane), PEI (polyethyleneimine), PAA (polyacrylic acid), PPz (polyphosphazene) among others. The TFCs cause the membranes to be quite strong, as the membranes need to be used and reused over again.

Ceramic membranes, which are much stronger than polymer membranes, are commonly made of Al, Si, Ti or Zr oxides and mixed oxides. The top layers of the membrane are prepared by sol-gel synthesis which converts a colloidal or polymeric solution of inorganic precursors (e.g. alkoxides or salts) to a gelatinous product. Viscous binders are usually added prior to layering followed by controlled calcination and finally sintering to form the ceramic membrane. The process can be controlled to provide membranes with specific MWCs (molecular weight cut-offs) and which can sustain defined solvents, temperatures and pressures.

Unfortunately at present not many of either type of membranes are available commercially, but undoubtedly many more will appear on the market in the near future. Commercially available SRNF include the *Se/RO membranes* (Koch Membrane Systems, USA), *Starmem membranes* (Membrane Extraction Technologies, UK), *SolSep membranes* (SolSep, The Netherlands), and *Desal-5 and Desal-5-DK membranes* (GE/Osmomics, USA) designed for aqueous applications, as well as for SRNF.

SRNFs have found applications in food chemistry, catalysis, chiral separations, petrochemical industry and pharmaceutical manufacturing. The preparation, practical considerations, and theoretical transport mechanisms of SRNFs have been critically reviewed by P. Vanezande, L.E.M. Gevers and I.F.J. Vankelecom [*Chem Soc Rev* **37** 365 2008].

### NANO METALS AND METAL DERIVATIVES

**Antimony (III) oxide nanopowder ( $\text{Sb}_2\text{O}_3$ )** [1309-64-4] **M 291.5, bulk density 0.5-0.6g/ml.** It is available in particle size of  $<250\text{nm}$  (BET). It is used in paints, pigments, enamels, glasses and flame-proofing canvas. It is toxic and should not be inhaled.

**Antimony-tin oxide nanopowder** [ATO,  $\sim(\text{SnO}_2)_{0.9}(\text{Sb}_2\text{O}_5)_{0.1}$ ] **M 167.0.** It is available in particle size of  $<50\text{nm}$  and has the alloy composition of 89-93% tin(IV) oxide and 7-11% antimony(V) pentoxide. The surface area is  $74\text{ m}^2/\text{g}$ . It is toxic and should not be inhaled.

**Cobalt** [7440-48-4] M 58.9, m 1495°, b 2870°, d<sup>20</sup> 8.71, resistivity 6.24 μΩcm/20°, Brinell hardness 125.

**Cobalt nanoparticles** for use in *magnetic fluids* (MFs) are available commercially with an average particle size of 10-12nm. It is in the form of a black powder and is prepared by a *thermolysis process* using a strict formula. Thus to dicobalt octacarbonyl {Co<sub>2</sub>(CO)<sub>8</sub>, 17.1g, 100mmol, see [10210-68-1]} under flowing argon is added Al(C<sub>8</sub>H<sub>17</sub>)<sub>3</sub> (4.4ml, 10mmol) [Co:Al = 10.1] dissolved in toluene (300ml) in one portion, and the mixture is stirred and heated at 110° under a reflux condenser for 18 hours. The reaction is carried out in an efficient fume cupboard because **POISONOUS** carbon monoxide evolution occurs. The colour of the solution changes to dark brown and a black precipitate separates from the clear solution. The mixture is cooled to 20°, and still under argon a further amount of Al(C<sub>8</sub>H<sub>17</sub>)<sub>3</sub> (1.2ml, ~30% of the initial quantity) is added, the temperature is raised to 110° and is heated for 3 hours more. The mixture is cooled to 20°, and stirred further for 16 hours. Smooth oxidation is carried out by bubbling (use a capillary) air through the mixture for 6 hours. The product is stirred overnight and the precipitated Co is allowed to settle during 2 hours. The supernatant is decanted and the solid is isolated in *toluene wet form*, or by drying *in vacuo* to give an air stable *Co nano-powder* (5.2g); both of which can be handled under laboratory conditions. The particle size of the Co, as determined by TEM and HRTEM is ~10±1.1nm (see further, however). The particle size is sensitive to the thermolysis conditions and to the Al(alkyl)<sub>3</sub>. With alkyl = Me, Et or *n*-octyl (C<sub>8</sub>H<sub>17</sub>) and a Co:Al ratio of 10:1 as above, the nanoparticle sizes are 3-4, 5-8 or ~10nm respectively. However with a Al(C<sub>8</sub>H<sub>17</sub>)<sub>3</sub> and Co:Al ratio of 1:2 the particle size was always ~5.4nm.

**Peptisation of Co nanoparticles:** The Co particles are suspended in toluene and peptised by adding 2ml of KorantinSH, oleic acid, LP-4 (a fatty acid condensation polymer, or AOT (sodium dioctylsulfosuccinate) as surfactants and sonicated in an ultrasound bath for 10 minutes to give a clear dark brown Co-MF. After drying a sample *in vacuo* elemental analysis indicated the presence 67.8% of Co and 0.98% of Al. These MFs were found to be stable for more than 5 months [by X-ray absorption near edge structure (XANES) measurements] under air. It should be pointed out that if the Al(C<sub>8</sub>H<sub>17</sub>)<sub>3</sub> is washed away from the Co particles immediately after the synthesis, the magnetic properties are markedly reduced after exposure to air. The chemical nature of the analysed nanoparticles shows that the stability of the local electronic and geometric structure is significantly affected by the surfactant used. Peptised Co nanoparticles can also be dispersed in kerosene or mineral oil as precursors for MFs.

**Cobalt nanoparticles (surface modified with L-cysteine ethyl ester) ethanol wet** are available as ~10nm black powder (wet with EtOH) which is best used within 3 months. The material can be easily transferred into stable aqueous suspensions, and can be used as a starting material for surface modifications (e.g. dextran coating) or the preparation of magnetic polymer microspheres.

**Copper-zinc alloy, 55-60% Cu** [12682-85-8] Cu/Zn, M 64.3. It has the composition of Cu (60% by wt) and Zn (40% by wt). The **nanopowder, has particle size <150 nm (TEM), 37-41% Zn basis**. It is a macro-agglomeration of ultra-fine particles, with sizes of *ca* 70 nm and is commercially available. It has trace Mn 1.5-2.5%, Sn 1.5-2.0% and Pb 0.5-2.5%.

**Copper zinc iron oxide, 98.5% trace metal basis (zinc copper ferrite, copper zinc ferrite)** [66402-68-4] M 416.3 (CuZnFe<sub>4</sub>O<sub>4</sub>), d<sub>4</sub><sup>25</sup> 5.5. It is a nano powder with particle size <100nm (BET). The actual particle size is <50nm (XRD) and is commercially available.

**Gold nanoparticles stabilised by gum Arabic (GA-AuNP).** AuNPs have been widely used for a variety of purposes and water soluble composites with gum Arabic (GA) are commercially available. Gum Arabic [9000-01-5] from *Acacia* gum is a branched polymer of galactose, rhamnose, arabinose and glucuronic acid as the Ca, Mg and K salts with M ~250,000, and is a non-toxic and physiologically compatible substance. GA-AuNP has been developed by Kannan, Katti and coworkers [Kattumiri et al. *Small* 3 333 2007]. It is prepared simply by mixing NaAuCl<sub>4</sub> with an aqueous solution of a non-toxic phosphine amino acid [P(CH<sub>2</sub>NHCHCH<sub>3</sub>COOH)<sub>3</sub> \*THPAL, see preparation below] reducing agent in the presence of a 0.2% aqueous solution of gum Arabic. This resulted in an immediate reaction which produced GA-labeled AuNPs in over 98% yield. The aqueous solutions are stable for months. Note that if GA is omitted, the reaction is also instantaneous but the solutions agglomerate within 4-6 hours. TEM images showed that the size of GA-AuNPs thus produced is 15-20nm. It is stable in the presence of cysteine, bovine and human serum albumins and does not agglomerate in 25% NaCl

solution. Formulations of it can be readily administered site specifically (*intravenously*), for diagnostic imaging (*computer tomography* CT), and is biocompatible for therapeutic applications in nanomedicine. GA-AuNPs are also suitable for nanoelectronics which include sensor design, MEMS applications, spin coating, self-assembly and formation of monolayers.

K.V. Katti and coworkers [Raghuraman et al. *J Am Chem Soc* **125** 6955 2003] prepared the reducing agent *\*tris[N-(R or S)-alaninylmethylene]phosphine* (\*THPAL) by adding tris(hydroxymethyl)phosphine (0.50g, 4.03mmol, see [2767-80-8]) in H<sub>2</sub>O (5ml) dropwise to R- or S- alanine (1.08g, 12mmol) in H<sub>2</sub>O (10ml) at 25° with stirring under N<sub>2</sub> for 1 hour. Evaporation *in vacuo* gave a white solid which, after washing with MeOH and drying *in vacuo* was pure product (90% yield). The <sup>1</sup>H NMR (300MHz, D<sub>2</sub>O) has δ at 1.38 (d, *J* = 6.0Hz, 9H, NCH(CH<sub>3</sub>)COOH), 3.47 (d, 6H, PCH<sub>2</sub>), 3.65 (m, 3H, NCH(CH<sub>3</sub>)COOH); the <sup>13</sup>C NMR (75MHz, D<sub>2</sub>O) has δ at 14.84 (s, NCH(CH<sub>3</sub>)COOH), 42.65 (d, *J*<sub>p-c</sub> = 12.82Hz, PCH<sub>2</sub>), 59.35 (d, *J*<sub>p-c</sub> = 5.77 Hz, NCH(CH<sub>3</sub>)COOH), 174.10 (s, NCH(CH<sub>3</sub>)COOH); and the <sup>31</sup>P NMR (121.5MHz, D<sub>2</sub>O) has δ at -39.9 (s); and the ESI-MS has calculated for C<sub>12</sub>H<sub>24</sub>N<sub>3</sub>O<sub>6</sub>P 337.3, found *m/z* 337.6. The X-ray crystal structure analysis of the S-enantiomer (4 H<sub>2</sub>O) revealed that the alanine portions were packed to form two-dimensional bilayers running parallel to (001). [See also Kannan et al. *J Am Chem Soc* **128** 11342 2006.]

**Gold nanoparticles stabilised by agarose (A-AuNP).** K.V. Katti and coworkers [Kattumuri et al. *Appl Phys Lett* **88** 153114 2006] prepared this similarly, but replacing GA by agarose. A-AuNPs were used for surface-enhanced Raman spectroscopic detection of DNA nucleosides as well as the other applications stated above.

**Gold (~1%) nanoparticles on titanium oxide extrudates (AUROLite Au/TiO<sub>2</sub>), on aluminium oxide extrudates (Au/Al<sub>2</sub>O<sub>3</sub>), on zinc oxide granulates (Au/ZnO), and on manganese oxide and carbon (Au/MnO<sub>x</sub>/C).** These are available commercially and used to catalyse some reactions. A 1% AuNP (~2nm) preparation on TiO<sub>2</sub> is used for catalytic purposes in 'Green' processes, e.g. in the presence of NaOMe in alcohols as solvents it catalyses the oxidation of aldehydes to esters using air as oxidant at ambient or lower temperatures [Marsden et al. *Green Chem* **10** 168 2008]. It also catalyses the oxidation of CO and H<sub>2</sub> using O or NO as oxidants [Walther et al. *J Catal* **260** 86 2008]. Au/Al<sub>2</sub>O<sub>3</sub>, Au/ZnO and Au/Mn<sub>x</sub>O/C are active catalysts for low temperature oxidation of CO and/or of CH<sub>4</sub> [Grisel & Nieuwenhuys *Catal Today* **64** 69 2001, Ma et al. *J Catal* **252** 119 2007, see also Freund *Catal Today* **117** 6 2006]. The preparations, properties and identification of these and related 'Gold on support' substances will be found in these references; and in Hugon et al. *Gold Bull* **41** 127 2008, Supansomboon et al. *Gold Bull* **41** 296 2008, and Steyn et al. *Gold Bull* **41** 318 2008.

**Indium-tin oxide nanopowder (ITO) [50926-11-9] M 264.9, surface resistivity of 100Ωcm.** Available in particle size <50nm. It is coated on glass for making transparent electrodes for screens (e.g. touch screens). However, graphene electrodes have better transparency and are tougher (see above) [Ahn et al. *Nature Nanotechnol* **5** 574 2010, DOI:10.1038/nnano.2010.132]

**Pd<sup>(0)</sup> EnCat™** is obtained by reducing the \*EnPd(OAc)<sub>2</sub> (0.4g, 0.4mmol/g, see below) in Et<sub>2</sub>O (8ml) with formic acid (8ml) at 45° for 2 hours. The mixture is cooled and the solid microcapsules are filtered through a polyethylene frit (20 micron porosity), washed with distilled H<sub>2</sub>O (3 x 30ml), Me<sub>2</sub>CO (3 x 30ml) and Et<sub>2</sub>O (3 x 30ml). The microcapsules are dried (at <0.5mm) at 45° for 5 hours to provide black **Pd<sup>(0)</sup> EnCat** microcapsules (0.35g, containing Pd nanoparticles of ~2nm size). It is an efficient catalyst for reductive ring-opening of epoxides to the respective alcohols which occurs in ~99% yields. The reaction requires either HCOOH/Et<sub>3</sub>N or H<sub>2</sub> as hydrogen donors, and the catalyst can be recycled ~10 times [Ley et al. *Org Lett* **5** 4665 2003].

**\*EnPd(OAc)<sub>2</sub> [Pd EnCat™]** is prepared from polymethylene polyphenylene di-isocyanate (SUPRASEC 5025, average functionality of 2.7) and Pd(OAc)<sub>2</sub> dissolved in CHCl<sub>3</sub> and then dispersed at 800rpm for 1 minute into an aqueous solution of Na lignosulfonate (Reax 100M), polyvinyl alcohol (Gohsenol GL 03) and the polyoxyethylene ether of butanol (Tergitol XD), using a mechanical overhead stirrer with a rotary blade. The resulting emulsion would have a particle size range of 20-250microns and the mixture is then gently shaken for 16 hours. The solid polyurea microcapsules formed are collected on a polyethylene frit (20 micron porosity), washed with de-ionised H<sub>2</sub>O (5 x 50ml), Me<sub>2</sub>CO (5 x 50ml), Et<sub>2</sub>O (3 x 50ml) and dried at room temperature.

Encapsulated Pd catalysts such as **Pd EnCat™** are available commercially. These particles are defined by their matrix content e.g. 30% or 40%, the latter having the smaller pore size. [See also Bremeyer et al. *Synlett* 1843 2002, Yu et al. *J C S Chem Commun* 678 2003, Vickerstaffe et al. *Org Biomol Chem* **1** 2419 2003.]

**Platinum nanoparticles on spherical polyelectrolyte brushes (SPB).** Pt crystals (2-3nm) were deposited on SPB using the procedure of M. Ballauff and coworkers [Schrinner et al. *Science* **323** 617 2009]. The SPB consists of a polystyrene core (~100nm diameter) onto which a cationic polyelectrolyte ( $^+H_3NCH_2CH_2OOC^-$ )<sub>n</sub> is attached [Schrinner et al. *Adv Mater* **20** 1928 2008]. A mixture of  $AuCl_4^-$  and  $PtCl_6^{2-}$  was reduced with  $NaBH_4$  on the surface of the cationic SPB to form alloy nanoparticles ( $Au_{45}Pt_{55}$  ratio) on the surface. The Au was removed with cyanide and  $O_2$ . Thus a solution of  $1.9 \times 10^{-5}$  M NaCN (4ml) was added dropwise within 25 minutes to a stirred suspension of AuPt-SPB (0.04 weight %) at room temperature under air. The high dilution of the cyanide is crucial for avoiding complete dissolution of the NPs and coagulation. Air was bubbled through the solution to remove completely the Au atoms. After 3 hours the colour of the solution was blue due to the formation of pure Pt NPs. [Schrinner et al. *Macromol Chem Phys* **208** 1542 2007.] Detailed examination of the composite particles by cryo-TEM, wide angle XRS and HR-TEM showed that the Pt NPs are well-defined faceted single crystals embedded in the SPB-chain layer. This composite system has excellent colloidal stability, and high catalytic activity with turn over numbers of as high as  $1580 \pm 50$  for the reduction of *p*-nitrophenol to *p*-aminophenol using  $NaBH_4$ , which is among the highest ever observed for this reaction. [For a review see H. Birch RSC *Chemistry World* **6** (3) 29 2009.]

**Rhodium nanoparticles entrapped in Rh/AlO(OH) matrix (~5wt% Rh loading).** The catalyst is prepared from a mixture of  $RhCl_3 \cdot xH_2O$ ,  $Al(secBuO)_3$  and 2-butanol (compare with Ru/AlO(OH) below) at  $100^\circ$  for 3 hours, aging in air for 1 day, filtering, washing the solid with  $Me_2CO$  and drying in air at  $25^\circ$ . It is estimated by high resolution transmission electron microscopy (HRTEM) and X-ray diffraction that the Rh particle sizes in the matrix are 2.5-3.0nm. It is an efficient catalyst for the hydrogenation of arenes, e.g. benzene, naphthalene, quinoline, at low pressures of  $H_2$  (~1 atmosphere) and room temperature in yields approaching 100% and with high turn over, with *ca* 1 mol% of Rh in the catalyst. [Park et al. *Chem Commun* 5667 2005].

**Ruthenium (nanoparticles) in aluminium oxide/hydroxide (~2.5wt% Ru loading).** In this catalyst, Ru is encapsulated in an aluminium oxy-hydroxide matrix where the metal has been shown by energy dispersive X-ray analysis (EDX) and X-ray photoelectron spectroscopy (XPS) to exist mainly as Rh(0) in the matrix. It is prepared in a one-pot synthesis through nanoparticle generation and gelation. Thus  $RuCl_3 \cdot xH_2O$  (52mg, 0.25mmol) and  $Al(secBuO)_3$  (2.5g, 10mmol) in EtOH (1.2ml, 20mmol) under a reflux condenser are heated, with stirring, at  $100^\circ$  for 1 hour, forming a black suspension to which  $H_2O$  (2.0ml) is rapidly added. The mixture is stirred for a further 30 minutes at  $100^\circ$ , filtered, the black solid is washed with  $Me_2CO$  and dried in air at  $25^\circ$  to give the catalyst as a grey powder (0.74g, 2.5wt% of Ru). It is used as a recyclable catalyst for the efficient oxidant-free alcohol dehydrogenation see Chapter 6, Catalysts-Part 1. [Kim, Park and Park *Org Lett* **8** 2543 2006.]

**Ruthenium nanoparticles stabilised in polymer ligands.** A general procedure was used for making these colloids (agglomerates) which involved introducing ruthenium (1,5-cyclooctadiene)(1,3,5-cyclooctatetraene) (158mg, 0.5mmol,  $[Ru(COD)(COT)]$  see [127382-91-6]) in a Fischer-Parr bottle and left *in vacuo* for 30 minutes, then THF (25ml, degassed by freeze-pump cycles) is added and the yellow solution is cooled to  $-80^\circ$ , after which the ligand (0.1mmol) in THF (~60ml) is introduced to the bottle. The bottle is pressurised under 3 bar of  $H_2$  and allowed to warm to room temperature. After ~20 hours a brown-black suspension was obtained. Complete decomposition of the Ru complex was checked by eliminating the  $H_2$  (e.g. by blowing  $N_2$  through) and taking a 3ml aliquot through a small  $Al_2O_3$  column when the filtrate should be colourless. The volume of the solution was then reduced to *ca* 15ml, pentane (50ml) was added, the resulting mixture was cooled to  $-80^\circ$  at which temperature a brown precipitate was formed. This was filtered off (or centrifuged), washed with pentane and dried *in vacuo*. The ligands used were PVP (poly vinylpyrrolidone), *n*-alkyl (C8, C12 and C16) amines and *n*-alkyl (C8, C12 and C16) thiols in varying amounts (from 0.2 to 1 equivalent) giving Ru percentages of ~55 to 80%. The colloids were characterised by microanalysis, IR after CO absorption, high resolution EM and wide-angle X-ray scattering. In THF and in a polymer matrix (Ru/polymer ~5%) crystalline particles of uniform mean size (1.1nm) and agglomerated particles (1.7nm) were obtained in PVP and cellulose acetate. The reaction with various concentration of alkyl amines and alkyl thiols lead to agglomerate particles or particles dispersed in solution both displaying mean size of 2-3nm. In the case with the amine ligands the particles were generally elongated with the tendency of forming worm or rod like structures at high amine concentrations, whereas  $^{13}C$  NMR studies showed that the stabilised particles were not fluxional, but thiol groups were oxidised to disulfide groups probably by the Ru surface. Because of their chemically reactive groups they are potentially very useful

in nanotechnology. [Pan et al. *J Am Chem Soc* **123** 7584 2001]. In a similar synthesis ruthenium nanoparticles were stabilised by organosilane fragments, by using *n*-octylsilane, giving particles of ~2.3nm diameter which is a very narrow distribution of size. They were characterised by TEM, as well as solid state <sup>13</sup>C MAS NMR. [Pelzer et al. *Chem Mater* **16** 4937 2004.]

**Silver acetate** [563-63-3] see also entry in Chapter 6, "Catalysis—Part I". Among the current intense activities in the synthesis of hybrid materials, where an inorganic phase is present as nanometer-sized particles dispersed throughout an organic matrix, is the preparation of a *ca* 190nm uniform film of exceptionally reflective and surface conductive silver "mirror" on a flexible poly(amic) matrix. The Ag-containing resin solutions are made first by dissolving AgOAc in a small volume of dimethylacetamide (DMAc, [127-19-5]) by adding 1.35 equivalents of 1,1,1-trifluoro-2,4-pentanedione (TFAH). As AgOAc is not very soluble in DMAc, addition of TFAH converts the acetate salt into the much more soluble AgTFAH complex. The poly(amic) matrix is prepared by co-polymerising 3,3',4,4'-benzophenonetetracarboxylic anhydride (BTDA, [2421-28-5]) with 4,4'-oxydianiline (APE, 4,4'-diaminodiphenyl ether, [101-80-4]), and making a 15%w/w solution in DMAc by stirring for 5 hours to give a viscosity of 1.7-1.8 dL/g at 35°. The poly(amic) solution is added to the Ag solution to give the desired Ag to BTDA/APE polymer ratio (10.7 to 13.0% Ag), cast into films (onto soda lime glass plates) to give cured films of 20-25 mm thickness. Air (10% relative humidity) is slowly flowed over the film for 18 hours, then thermally cured in a forced air oven with a cycle involving heating at 135° for 20 minutes, holding for 1 hour, heating at 300° for 4 hours, and holding at 300° for 7 hours. This process reduces Ag(I) to the metal and releases all the volatiles. Uniform nanometer thin films of Ag metal (see above) are thus formed in the polymer matrix. Characterisation by X-ray diffraction, transmission (TEM) and scanning (SEM) microscopy, tapping mode atomic force microscopy, X-ray photoelectron microscopy (XPS), and conductivity, reflectivity, thermal and mechanical measurements showed that the films have excellent properties, and are stable in air below 325°. This 'film-to-film' method minimises the amount of Ag used. The applications for this Ag-copolymer silvered films with exceptional reflectivity, thermal, and surface conductivity are immense, and include the fabrication of lightweight optical mirrors, sunshields (e.g. in space telescopes), radiofrequency antennas for management of electromagnetic signals in space, Ag coated tubing to deter catheter-induced urethritis, elastomeric devices, concentration of solar energy, flexible conductive patterned surfaces and tapes. [Southward et al. *Chem Mater* **11** 501 1999.]

**Silver nanoparticles on spherical electrolyte brushes (SPB).** Compare these with the above Pt NPs on SPB. They were also developed by M. Ballauff and coworkers [Lu et al. *J Phys Chem C* **111** 7676 2007]. In this case Ag nanoparticles were synthesised onto the polystyrene (PS) core—poly(acrylic acid) (PAA) SPB particles *in situ* by photoemulsion polymerisation. The PS core with a thin layer of photo initiator {2-[*p*-(2-hydroxy-2-methylpropiophenone)]-ethylene glycol-methacrylate} was made by standard emulsion polymerisation procedures also developed by M. Ballauff and coworkers [Gua et al. *Macromolecules* **32** 6043 1999]. The PS-PAA-Ag composite particles were prepared *in situ* by mixing diluted PS core solution (1 wt %) with a defined quantity of functional monomer silver acrylate (30 mol% of the amount of styrene), degassing and refilling with N<sub>2</sub> (5 times). Then the mixture was irradiated under N<sub>2</sub> with UV-VIS radiation at ~25° for 90 minutes while being vigorously stirred to keep the mixture homogeneous. Coagulum was removed by filtration through glass wool and the Ag nanocomposite particles were washed by dialysis against purified H<sub>2</sub>O using a cellulose nitrate membrane (100nm pore size). DLS showed that the polyelectrolyte brushes were on the PS core; TEM, Cryo-TEM, and wide-angle XRS revealed that the Ag NPs in the brushes were crystalline with a diameter of 3 ± 1.2nm.

This Ag-composite is stable and has very good catalytic activity. Reduction of *p*-nitrophenol using NaBH<sub>4</sub> is very fast, albeit slightly slower than with Pt or Pd immobilised on related carriers.

**Silver nanoparticles [AgNP], on gum Arabic, gelatin or starch.** Available as yellow-brownish liquids. They are water soluble in the form of 5-10nm spheres (10-15nm for starch) and have UV with λ<sub>max</sub> at 405-410nm. The gum Arabic (GA), gelatin and starch stabilise these particles and can be used for sensor design applications, for *in vitro* as well as *in vivo* antimicrobial and antifungal purposes.

**Silver-tin alloy nanoparticles, M 118.3.** These particles are available in <150nm in size and contain 3.5% Ag with the formula Ag<sub>0.035</sub>Sn<sub>0.965</sub>. They have a surface area >5.4 m<sup>2</sup>/g.

**Stannic oxide nanopowder (Sn(IV)O<sub>2</sub>)** [18582-10-5] **M 150.7**. It is available in particle size of <100nm. It is used for polishing metal or glass, used in coloured glass, enamels and as a mordant for colouring fabrics. It is toxic and should not be inhaled.

**Zinc nano metallic powder (ZnNP)** [7440-66-6] **M 65.4**. ZnNP is available as a gray powder of spherical particles of 75-125nm size. It is a very reactive form of the metal as it has a high surface area. The powder is particularly reactive with organo-halides to form organo-Zn compounds and in azo-coupling reactions. Sonication will de-agglomerise the nano Zn into particles of 35nm size. ZnNP with particle size <50nm is also available commercially with BET surface area of 35-50 (m<sup>2</sup>/g).

## SOME ORGANIC AND METAL-ORGANIC COMPOUNDS USED IN NANOTECHNOLOGY

**1-Adamantane thiol (tricyclo[3.3.1.1<sup>3,7</sup>]decane-1-thiol)** [34301-54-7] **M 168.3, m 95-97<sup>o</sup>, 100-102<sup>o</sup>, 99-106<sup>o</sup>, pK ~11.2**. This thiol was prepared by various methods namely by formation of 1-adamantylisothiourinium bromide from 1-adamanyl bromide followed by alkaline hydrolysis [Geigy A.-G. **Belg. Pat** 629,370 1963, *Chem Abstr* **60** 9167 1963], by refluxing 1-adamantanol with Lawesson's reagent (see [19172-47-5]) in toluene [Nishio *J Chem Soc, Perkin Trans I* 1113 1993], and by photolysis of *N*-(1-adamanylcarbonyloxy)pyridin-2-thione [Barton reaction: Barton et al. *Tetrahedron Lett* **35** 6057 1994]. The thiol is generally purified by drying over P<sub>2</sub>O<sub>5</sub>, dissolving in Skellysolve B (see Aliphatic Compounds, Chapter 4) or *n*-hexane, passing through a column of Alumina in *n*-hexane, and eluting with *n*-hexane/EtOAc (97:3). Store it dry under N<sub>2</sub> or argon. The IR (CCl<sub>4</sub>) has  $\nu_{\max}$  at 2190, 2850, 1450 and 2565 (S-H) cm<sup>-1</sup>, the <sup>1</sup>H NMR (CCl<sub>4</sub>) has  $\tau$  at 8.06 (s, 9,  $\beta$  and  $\gamma$  H), 8.29 (s, 6,  $\delta$  H) and 8.57 (s, 1 S-H); and the MS has *m/z* (peak height) 168 (14, M<sup>+</sup>), 136 (12), 135 (100), 93 (15), 79 (17). The *S*-2,4-dinitrophenylsulfenyl derivative gives yellow needles from MeOH with **m 159-160<sup>o</sup>**. [Tanner & Brownlee *J Canad Chem* **51** 3366 1973]. The *S*-CH<sub>2</sub>CO<sub>2</sub>Et derivative (from NaNH<sub>2</sub>/BrCH<sub>2</sub>CO<sub>2</sub>Et) has **b 117-118<sup>o</sup>/0.001mm**, and on hydrolysis (NaOH/EtOH, 5 hours) gave  $\alpha$ -(adamant-1-ylthio)acetic acid **m 68-70<sup>o</sup>** (hexane/pentane). Similarly obtained are  $\alpha$ -(adamant-1-ylthio)propionic acid **m 142-144<sup>o</sup>** (cyclohexane),  $\alpha$ -(adamant-1-ylthio)butyric acid **m 113-114<sup>o</sup>**,  $\alpha$ -(adamant-1-ylthio)isovaleric acid **m 134-145<sup>o</sup>** (cyclohexane),  $\alpha$ -(adamant-1-ylthio)caproic acid **m 74-76<sup>o</sup>**,  $\alpha$ -(adamant-1-ylthio)- $\alpha$ -phenylacetic acid **m 122-124<sup>o</sup>** (cyclohexane/hexane) as well as other acids that are reported [Geigy A.-G. **Belg. Pat** 629,370 1963, *Chem Abstr* **60** 9167 1963]. 1-Adamantane-thioether is a good novel sialoside protecting group which is also used for coupling with adequately protected sugars [Critch & Li *J Org Chem* **72** 7794 2007].

1-Adamantanethiolate molecules have been used in a new patterning technique (microdisplacement printing) to form self-assembled mono-layers (SAMs) on an Au(111) surface for contact printing. The attached thiolate molecular film hinders lateral surface diffusion of the patterning molecules allowing the use of such molecules which are too mobile to pattern by other methods. It is used for SAMs Cat 659452-5G molecular spacer in nanotechnology [ *J Org Chem* **72** 7794 2007, Dameron et al. *Nano Lett* **5** 1834 2005].

**Coronene (hexabenzobenzene, [6]circulene)** [191-07-1] **M 300.4, m 420<sup>o</sup>, 425-428<sup>o</sup>, 428<sup>o</sup>, 438-440<sup>o</sup>, 440<sup>o</sup> (uncorrected), 442<sup>o</sup> (sealed tube) b 525<sup>o</sup>/atm**. Purify coronene by passing a solution in \*C<sub>6</sub>H<sub>6</sub> through a short Al<sub>2</sub>O<sub>3</sub> column and eluting with \*C<sub>6</sub>H<sub>6</sub>, followed by evaporation and recrystallisation from \*C<sub>6</sub>H<sub>6</sub> or xylene (**m 245-248<sup>o</sup>**), and sublimation at 200-300<sup>o</sup>/20mm, or 420<sup>o</sup>/atm requiring ~2 hours. Crystallisation from \*C<sub>6</sub>H<sub>6</sub> yields amber-yellow needles (5—15mm long). Alternatively, coronene in \*C<sub>6</sub>H<sub>6</sub> is added to a saturated solution of picric acid in EtOH whereby the *picrate* (**m 310-325<sup>o</sup> dec**) separates as red needles. The *picrate* in \*C<sub>6</sub>H<sub>6</sub> is passed through an Al<sub>2</sub>O<sub>3</sub> column and eluted with \*C<sub>6</sub>H<sub>6</sub> when picric acid remains strongly adsorbed on the column. The eluate is evaporated and the residue recrystallised from toluene, and sublimed (yellow needles) at 250<sup>o</sup>/high vacuum. It has been distilled at atmospheric pressure using a naked flame. In concentrated H<sub>2</sub>SO<sub>4</sub> a green-yellow solution is obtained. The *picrate* forms bright red needles from \*C<sub>6</sub>H<sub>6</sub> with **m 310-325<sup>o</sup>**. The 1,3,5-trinitrobenzene adduct gives orange needles which decompose at high temperature. The UV spectrum in \*C<sub>6</sub>H<sub>6</sub> contains 16 sharp peaks between 290 and 428nm, but in EtOH it has  $\lambda_{\max}$  (log  $\epsilon$ ) at 228 (4.08), 252 (3.88) and 290 (5.10) nm. [Baker et al. *J Chem Soc* 1114 1951, Baker et al. *J Chem Soc* 1120 1951, Clar & Zander *J Chem Soc* 4616 1957, Hopff & Schweizer *Helv Chim Acta* **42** 2315 1959, *Beilstein* **5** III 2661, **5** IV 2830.] It is a nano-channel organic semiconductor [Newman et al. *Chem Mater* **16** 4436 2004] and is used to synthesise MBE-grown layered superconductors [*Japanese J Appl Phys, Part 1* **34** 3837-3845 1995].

**6-Mercaptohexylferrocene** [134029-92-8] **M 302.3, b 321-353°**. This thiol is prepared in four steps from ferrocene. Thus a mixture of 6-bromohexanoyl chloride (3.2g, 15mmol, see [22809-37-6]) and anhydrous  $\text{AlCl}_3$  (2.0g, 15mmol) in cold ( $0^\circ$ )  $\text{CH}_2\text{Cl}_2$  (40ml), which is stirred (20 minutes) under argon, is added dropwise with stirring to ferrocene (2.79g, 15mmol) in  $\text{CH}_2\text{Cl}_2$  (40ml) over a period of 20 minutes when the colour of the solution turns purple. After stirring for 2 hours, the stirring rate is increase while  $\text{H}_2\text{O}$  (20ml) is added slowly, and the solvent boils gently. The organic layer is collected after 10 minutes, washed with  $\text{H}_2\text{O}$  until the washings are neutral, dried ( $\text{MgSO}_4$ ), filtered, concentrated *in vacuo*, and the residue is subjected to chromatography on silica-gel (230-400 mesh) with hexane/EtOAc (100:1) as eluent. The second orange band (the first is ferrocene) is collected and gives *6-(ferrocenylcarbonyl)pentyl bromide* (4.5g, 85%) on evaporation. A part of it (1.61g, 7.4mmol) and Zn/Hg amalgam (freshly prepared from 7.4g of granulated Zn and 0.5g of  $\text{HgCl}_2$ ) in  $\text{H}_2\text{O}$  (25ml), 12M HCl (20ml) and toluene (30ml) are stirred vigorously so that the two liquid phases and the solids are in close contact, and are heated under reflux for 16 hours during which time 12M HCl (2 x 5ml) is added. The mixture is cooled to  $\sim 25^\circ$ , the organic layer is collected washed with  $\text{H}_2\text{O}$ , dried ( $\text{MgSO}_4$ ), filtered, concentrated, and the residue is purified through a silica-gel column as before. The yellow-brown band provides the required *6-ferrocenylhexyl bromide* (2.33g, 90%). The bromide (1.11g, 3.2mmol) and thiourea (0.24g, 3.5mmol) in absolute EtOH (20ml) is then boiled under reflux in an atmosphere of argon. After boiling for 18 hours, the solvent is removed *in vacuo*, treated with aqueous KOH (0.20g, 3.5mmol, in 20ml of  $\text{H}_2\text{O}$ ) and boiled under reflux in an atmosphere of argon for 2 hours. The orange-yellow oil that settles on cooling is extracted into  $\text{Et}_2\text{O}$  (3 x 50ml), the combined extracts are dried ( $\text{MgSO}_4$ ), concentrated, and chromatographed as before. The first yellow band gives the desired **6-ferrocenylhexanethiol** (0.68g, 62%) on evaporation. It is identified by its  $^1\text{H}$  NMR (300 MHz,  $\text{CDCl}_3$ ) which has  $\delta$  at 1.35 (t, 1H, SH), 1.39 (br, 4H,  $\text{FcCH}_2\text{CH}_2\text{CH}_2$ ), 1.52 (q, 2H,  $\text{CH}_2\text{CH}_2\text{CH}_2\text{SH}$ ), 1.63 (q, 2H,  $\text{CH}_2\text{CH}_2\text{SH}$ ), 2.34 (t, 2H,  $\text{FcCH}_2$ ), 2.53 (q, 2H,  $\text{CH}_2\text{SH}$ ), 4.06 (t, 2H,  $\text{C}_5\text{H}_4$ ), 4.07 (t, 2H,  $\text{C}_5\text{H}_4$ ), 4.11 (s, 5H,  $\text{C}_5\text{H}_5$ ); and EI-MS  $m/z$  at 302 ( $\text{M}^+$ ). The second yellow band contains the corresponding *disulfide* which can be reduced and used. [Creager & Rowe *J Electroanal Chem (Lausanne Switz)* **370** 203 1994, cf Yu et al. *J Org Chem* **66** 2937 2001.] This and other related ferrocenyl-alkane thiols have been synthesised and characterised electrochemically in solution and in self-assembled monolayer (SAM) films on gold electrodes. The affinities of these on the electrode surface have been described in detail. These red-ox active thiols can be used for selective electrochemical sensing and preparing a new type of thiol self-assembled sensing devices (SAMs). [Creager & Rowe *J Electroanal Chem (Lausanne Switz)* **370** 203 1994.] Gold nanoparticles covered fully with 6-ferrocenylhexanethiolate ligands, with average composition  $\text{Au}_{225}(\text{6-ferrocenylhexanethiolate})_{42}$ , were prepared, and had a unique combination of adsorption properties on Pt electrodes which were studied comprehensively. The adsorbed layers are so robust that sub, mono and multi nanoparticulate layers on the electrodes can be transferred to fresh electrolyte solutions in which they show stable ferrocene voltametry over long time periods. [Stiles et al. *J Am Chem Soc* **130** 1856 2008.]

**NPs for drug delivery** to tissues and particularly to brain, i.e. traverse the blood-brain-barrier, include conjugates of dextran, agarose, chitosan polyacetic acid (PLA), poly(lactic-co-glycolytic acid) (PLGA) and NPs with lipophilic moieties, e.g. polyethylene glycol (PEG) to which drugs can be bound have been explored by Fresta and co-workers [Celia et al. *Med Res Rev* **31** (3) 716-756 2011], and shown to be useful devices.

**Tetra-*n*-propylammonium perruthenate (TPAP, tetrapropyl tetraoxoruthenate)** [114615-82-6] **M 351.4, m 160°(dec)**. This oxidant (at 5 mol%) acts catalytically in oxidising 2-substituted *N*-1, *N*-3-dihydroxy-imidazolidines (1mol) in  $\text{CH}_2\text{Cl}_2$  at room temperature in 1-12 hours in the presence of *N*-methylmorpholine *N*-oxide (1.1mol) to provide the respective 2-substituted imidazoline-1-nitronyl-3-nitroxide radicals (NNRs) in 44-90% yields. NNRs have found applications in Physics, Chemistry and Nanotechnology for the development of organic molecular magnets [Palacio et al. *Phys Rev Lett* **79** 2336 1997] and single chain magnets [Caneschi et al. *Angew Chem Int Ed* **40** 1760 2001]. Gorini et al. *Synlett* 948 2006; for further reading see TPAP entries on p. 656 in Chapter 5, 'Metal-Organic Compounds', and on p. 708 Chapter 6, 'Catalysis-Part 1'.

## SAFETY ISSUES

Materials of nanometer size have been present on earth almost since its formation  $\sim 4.6$  billion (eon) years ago, as well as in the universe. Recently J. Cami, and coworkers have discovered large clouds in the planetary nebula Tc1 in which  $\text{C}_{60}$  and  $\text{C}_{70}$  fullerenes were identified by infra red spectroscopy. [*Science* **329** (Issue 5996) 1180

2010, reviewed by A. Extnance RSC *Chemistry World* 7 (9) 21 2010, and L. Howes RSC *Chemistry World* 8 (1) 34 2011.] During the past 30 years or so the development of nanomaterials and their wide ranging applications have avalanched, and progress is continuing. The safety aspects of nanomaterials have been of concern for many years, and it is generally accepted that conscious efforts have to be made to protect the individual workers in the field, the general public and the environment. Research on safety aspects is slowly gaining pace, and it has already been shown in mice that carbon nanotubes (CNTs) exhibit asbestos-like pathogenicity when inhaled. [K. Donaldson and coworkers: Poland et al. *Nature Nanotechnol* 3 423 2008, J.C. Bonner and coworkers: Ryman-Rasmussen et al. *Nature Nanotechnology* 4 747 2009, Reviewed by H. Birch RSC *Chemistry World* 6 (12) 24 2009.] The movement of nanomaterial from lungs to lymph nodes to the blood stream and clearance through the kidneys has been demonstrated by J.V. Frangioni and A. Tsuda and their coworkers [Choi et al. *Nature Biotechnol* 28 1300 2010, reviewed by L. Howes RSC *Chemistry World* 7 (12) 24 2010.] More recently W. Wohlleben and coworkers have subjected two thermoplastic materials and two cement materials infused with different nanoparticles (i.e. nanotubes and nanoparticle in polymers filters and in inorganic matrices) to gentle abrasion, high speed sanding and UV radiation. They did not find significant release of nanopowders into the atmosphere in which the sizes of the nanoparticles were very different from those in control powders. Moreover, rats exposed to the dust fared no worse than those exposed to dust that contained nanocomponents. However, further studies are in progress. [Wohlleben et al. *Small* 7 (16) 2384 2011, DOI: 10: 1002/smll.201002054; reported by K. McAlpine *Chemistry World* 8 (8) 28 2011.] Safety was rather slack in the early days, but it is being realised that nanomaterials should be treated differently from hazardous chemicals due to their unknown long-term effects. Nano-enabled hoods and nano-enabled-gear for personnel are being used, but only to a small extent. Safety protocols have been written and updated, but nano-regulation is only slowly creeping in [see special report by V. Gill RSC *Chemistry World* 6 (4) 10 2009.] A comprehensive report on nanosafety practices in research laboratories worldwide was compiled by J. Santamaria and coworkers recently, where they detail an up-to-date account of the present situation regarding to safety issues. [Balas et al. *Nature Nanotechnol* 5 93 2010; Reviewed by K. McAlpine RSC *Chemistry World* 7 (3) 15 2010.]

***Our advice is that in the manipulation and disposal of nanomaterials the same guidelines for asbestos handling should be used. In addition, further precautions should be exercised which should consider the known toxicity and the chemical nature of the nanomaterials in use.***

The following is a list of more recent books that were published on legal, personal, industrial, and environmental safety aspects of nanotechnology:

L.L. Bergeson (Ed), *Nanotechnology: Environmental Law, Policy and Business Considerations*, American Bar Association, ABA Publishing USA, 2010. ISBN 9781604225826.

J-Y. Botero, *Environmental Nanotechnology*, McGraw Hill, 2007. ISBN 97800714777505, 0071477500.

R.E. Hester and R.M. Harrison (Eds), *Nanotechnology: Consequences for Human Health and the Environment*, RSC Publ. 2009. ISBN 9781847559562, 1847559565.

C. Kumar (Ed), *Nanomaterials: Toxicity (Health and Environment Issues)*, WILEY-VCH, Weinheim, 2006. ISBN 9783527313853, 3527313850.

J.R. Lead and E. Smith (Eds), *Environmental and Human Aspects of Nanotechnology*, J. Wiley & Sons, NY, 2009. ISBN 9781405176347.

I. Linkov and J. Steevens, *Nanomaterials: Risks and Benefits*, Springer, 2008. ISBN 9781402094903 (PB), 9781402094907 (HB), 9781402094910 (e-book).

N.A. Montiero-Riviere and C.L. Tran, *Nanotoxicology: Characterisation, Dosing and Health*, Informa Healthcare Publ, 2007. ISBN 9781420024147, 1420045148.

S.C. Sahn and D.A. Casciano, *Nanotoxicity*, J. Wiley & Sons, NY, 2009. ISBN 9780470741375.

K.Sellers, *Nanotechnology and the Environment*, CRC Press, 2009. ISBN 9781420060195, 1420060198.

## BIBLIOGRAPHY

D. Amabilino (Ed), *Chemistry at the Nanoscale: nanoparticles, surfaces, materials and more* Wiley-VCH, Weinheim, 2009. ISBN 978352730134.

V.E. Borisenko, S.V. Gaponenko and V.S. Gurin (Eds), *Physics, chemistry and application of nanostructures: reviews and short notes: proceedings of the international conference, Nanomeeting 2007*. ISBN 978981275990, 981275996.

B. Bhushan (Ed), *Handbook of Nanotechnology 2<sup>nd</sup> edn*, Springer, Berlin, 2007. ISBN 10: 354029855-x, 13: 9783540298557.

K. Deffeyes and S. Deffeyes, *Nanoscale: visualizing an invisible world*, MIT Press, Cambridge USA, 2009. ISBN 9780262012836.

M.S. Dresselhaus, G. Dresselhaus and P.C. Eklund, *Science of Fullerenes and Carbon Nanotubes*, Academic Press, San Diego, 1996. ISBN

J.L. Feather and M.F. Azner, *Nanotechnology: Education and Workforce Development*, CRC Press, Boca Raton, FL, 2011. ISBN 9781420053944.

B.D. Feldman, *Materials Chemistry*, Springer, Dordrecht, 2008. ISBN 978-1-4020-6119-6, 978-1-4020-6120-2 (e-book).

J.H. Grassian, *Nanoscience and Nanotechnology*, J. Wiley & Sons, 2008. ISBN 978047081037.

J.A. Helsen and Y. Missirlis, *Biomaterials: Biological and Medical Physics, Biomedical Engineering*, Springer Verlag, 2010. ISBN 9783642125317, 9783642125324 (e-book).

C. Hess and R. Schlögl (Eds), *Nanostructured Catalysts*, RSC Publ. Nanoscience & Nanotechnology, 2011. ISBN 9780854041866.

G.L. Hornyak and H.F. Tibbals, *Fundamentals of Nanotechnology*, CRC Press, Boca Raton, FL, 2009. ISBN 978140048032, 1420048031.

G.L. Hornyak, H.F. Tibbals and J. Dutta, *Introduction to Nanoscience and Nanotechnology*, CRC Press, Boca Raton, FL, 2009. ISBN 9781420047790, 1420047795.

M.S. Johal, *Understanding Nanomaterials*, CRC Press, Boca Raton, FL, 2011. ISBN 9781420073102.

C.S.S.R. Kumar, *Nanocomposites*, Wiley-VCH, Weinheim, 2010. ISBN 9743527321681.

C.S.S.R. Kumar, *Nanostructured Thin Films and Surfaces*, Wiley-VCH, Weinheim, 2010. ISBN 9783527321551, 3527321551.

D.P. O'Mathuna, *Nanoethics: Big Issues with Small Technology*, Continuum International Publ. Co. London, 2009. ISBN 9781847063946 (HB), 9781847063953 (PB).

G.A. Ozin, A. Arsenault and L. Cademartiri, *Nanochemistry: A Chemical Approach to Nanomaterials*, RSC Publ., 2009. ISBN 9781847558954.

J. Ramsden, *Nanotechnology: An Introduction*, Elsevier Science and Technology, 2010. ISBN 9780080964478, 0080964478.

H-G. Rubahn, *Basis of Nanotechnology*, 3<sup>rd</sup> Edition, Wiley-VCH, Berlin, 2008. ISBN 13: 9783527408009, 10: 3527408002.

G. Schmid (Ed), *Nanotechnology: Principles and Fundamentals Vol 1* (ISBN 9783527317325); H. Krug (Ed), *Nanotechnology: Environmental Aspects Vol 2* (ISBN 9783527317356); R. Wasser (Ed), *Nanotechnology: Information Technology Vol 3* (ISBN: 9783527317387); R. Wasser (Ed), *Nanotechnology: Information Technology <pt.1> Vol 4* (ISBN: 9783527317370), Wiley-VCH, Weinheim, 2008.

G. Schmid (Ed), *Nanoparticles: From Theory to Application*, 2<sup>nd</sup> edition, Wiley-VCH, 2010, ISBN 3527325891 (available from the internet with Babylon Translation Software free download, 538 pp, 12 MB as PDF).

# GENERAL INDEX

## A

- Abbreviations 103, 555, 771
- Abderhalden pistol 29
- Abrin A & C 798
- L-Abrine (*N*-Methyltryptophan) 972
- Acac (Acetylacetone) 108
- Acacetin (5,7-Dihydroxy-4'-methoxyflavone) 431
- ACE (Angiotensin converting enzyme) 800
- ACES [*N*-(2-Acetamido)-2-aminoethanesulfonic acid] 104
- AcetamidoTEMPO (4-Acetamido-2,2,6,6-tetramethylpiperidine-1-oxyl) 397
- Acetazolamide (5-Acetamido-1,3,4-thiadiazole-2-sulfonamide) 397
- Acidomycin 900
- Acridine Orange (3,6-Bisdimethylamino)acridine 399
- Acridine Yellow G (2,7-Dimethylacridine-3,6-diamine hydrochloride) 400
- Acrinol (6,9-Diamino-2-ethoxyacridine *dl*-lactate monohydrate, Rivanol) 426, 814
- Actarit (*p*-Acetamidophenylacetic acid) 259
- Actidione (Cycloheximide) 424
- Actinomycin C 858
- Actinomycin D (Dactinomycin) 859
- Actinospectacin 897
- ADA [*N*-(2-Acetamido)iminodiacetic acid] 104
- Adam's catalyst 666, 667
- ADCC-linker (2-Acetyl-1-methyl-3,5-dioxo-1-methylcyclohexane carboxylic acid) 220
- Adermine (Vitamin B<sub>6</sub> HCl) 904
- Adkin's catalyst (Lazier catalyst, Copper chromite) 663
- ADP (Adenosine diphosphate) 859
- L-Adrenaline [L-(-)(3,4-Dihydroxyphenyl)-2-methylaminoethanol] 860
- Adrenor (L-Noradrenaline) 889
- Adsorbents, graded 48
- solvents 48
- Aerosol-OT (Sodium di(ethylhexyl)sulfosuccinate) 530
- Affinity chromatography, 26, 28, 60, 766, 768
- Agglutinin 799
- AIBN [2,2'-Azobis(isobutyronitrile)] 117, 915
- AICAR HCl (4-Aminoimidazole-5-carboxamide HCl) 404
- Ajmalicine (δ-Yohimbine, Raubacine) 401
- Ajmaline (γ-Yohimbine) 401
- ALA-HCl (δ-Aminolaevulinic acid HCl) 778
- Albumin 799
- Aldrithiol 2 & 4 (2,2'- & 4,4'-Dipyridyl disulfide) 436
- Aleuritic acid (*RS*-erythro-9, 10, 16-Trihydroxyhexadecanoic acid) 112
- Algin (Sodium alginate) 650
- Alkaline base bath 3
- n*-Alkylammonium chlorides 112, 754
- n*-Alkyltrimethylammonium bromides 112
- Allocholic acid (5α-Cholanic acid) 844
- Allocholesterol 847
- Allopregnane (5α-Pregnane) 853
- Aloin, Barbaloin 817
- Altracin 862
- Alumina, 30
- activation 22
- TLC 20
- preparation 21
- standardisation 21
- Aluminium amalgam 156
- Amanozine (2-Amino-4-anilino-triazine) 402
- Amantadine HCl (1-Adamantyl-Amine HCl) 221
- Amberlite IRA-904 261
- Ambucetamide [α-(*p*-methoxyphenyl)acetamide] 311
- t*-AMCHA (*t*-4-AminoethylCyclohexane-1-carboxylic Acid) 776
- Amechol (*O*-Acetyl-β-methylcholine chloride) 857
- AFM (Atomic Force Microscopy) 908, 909
- α-Amino acids 772, 775, 801
- Aminometradine [1-Allyl-6-amino-3-ethyluracil] 402
- Aminopentamide (Dimevamide) 758
- Aminopterin 861
- Amiphenazole (2,4-Diamino-5-phenylthiazole) 426
- Amitrol (3-Amino-1,2,4-triazole) 406
- Amminosidin (Paromomycin sulfate) 891
- Ammoidin (Xanthotoxin) 906
- 3'-AMP (Adenosine 3'-monophosphate) 859
- 5'-AMP (Adenosine 5'-monophosphate) 859
- Amphicol 865, 866
- Ananase (Bromelain) 800
- Anaprel 483
- Android (17α-Methyltestosterone) 853
- Aneurine HCl (Vitamin B<sub>1</sub>) 903
- Aneurine monophosphate HCl 900
- Angiotensin 800
- Angiotensin-converting enzyme ACE 800
- Angiotensinogen 800
- Anhydron (Dehydrite, MgClO<sub>4</sub>) 5, 31, 588
- Anion exchange resin (s) 24, 25, 48, 65, 861
- Aniracetam (1-[4-Methoxybenzyl]-2-pyrrolidinone) 861
- Anserine 779
- Anthrarufin (1,5-Dihydroxy-9,10-anthraquinone) 270
- Apiezon L, M and T 12, 47
- Apigenin (4',5,7-Trihydroxyflavone) 492
- R*-Apomorphine 862
- Apoterin (Rescinnamine) 483
- Aquamycin (Acetylenedicarboxamide) 109
- Aqua fortis (Nitric acid) 593
- Aqua regia 561, 578, 579, 585, 595, 598, 606
- Arabitol (D, L, DL-Arabinitol) 818
- Archin, Emodin (1,3,8-Trihydroxy-6-methyl-9,10-anthracenedione) 332
- Aristolene, [(+)-Calarene] 224
- Arterenol (L-Noradrenaline HCl) 889
- Ascarite (NaOH-Coated silica/asbestos adsorbent) 561, 592, 594
- Aspirin (Acetylsalicylic acid) 261
- Atebrine, Mepacrine (3-Chloro-9(4-diethylamino-1-methyl)butylamino-7-methoxy)acridine dihydrochloride) 481

Atom economy 92  
 Atophen, Cinchophen (2-Phenylquinoline-4-carboxylic acid) 472  
 ATP (Adenosine 5'-triphosphate) 860  
 Atrolactic acid (2-Hydroxy-2-phenyl-propionic acid) 270  
 Auramine O (4,4'-Bis-dimethyl-aminobenzophenone imine hydrochloride) 271  
 Aureolic acid (Plicamycin) 885  
 Aurin, corallin 379  
 Automated column chromatography 28  
 Avidin 800  
 Azaserine 780  
 Azurin 800

**B**

Bacitracin 862  
 BAL (2,3-Dimercaptopropan-1-ol, British Anti-Lewisite) 147  
 BAO [2,5-Bis(4-aminophenyl)-1,3,4-oxadiazole] 412  
 Barbital (5,5-Diethylbarbituric acid) 430  
 Barbitone acid Na salt (Sodium 5,5-diethylbarbiturate) 650  
 Baryta water 562  
 Bathocuproine (2,9-Dimethyl-4,7-diphenyl-1,10-phenanthroline) 432  
 9BBN (9-Borabicyclo[3.3.1]nonane) 503  
 BBOD (2,5-Bis(4-biphenyl)-1,3,4-oxadiazole) 412  
 BCDC (8*S*,9*R*-*N*-Benzylcinchonidinium chloride) 752  
 BCNC (8*S*,9*S*-*N*-Benzylcinchoninium chloride) 752  
 Bekanamycin (Kanamycin B) 882  
 Benzene, care in use of 103, 555, 771  
 purification 272  
 Benzocaine (Ethyl *p*-aminobenzoate) 333  
 S-Benzylisothiuronium chloride 281

BES [*N,N'*-Bis(2-hydroxyethyl)-2-aminoethane sulfonic acid] 118  
 Bestmann ylide [(Triphenylphosphoranylidene) ketene] 550  
 BET 909  
 Betaine (1-Carboxymethyl-*N,N,N*-trimethylammonium zwitterion) 780  
 Betol (2-Naphthyl salicylate) 357  
 BH<sub>4</sub> 2HCl 899  
 BHT (2,6-Di-*t*-butyl-*p*-cresol) 311  
 BICINE [*N,N*-Di-(2-hydroxyethyl)-glycine] 784  
 (+)-Bicuculine 863  
 BINAP [2,2'-Bis(diphenylphosphino)-1,1'-binaphthyl] 721, 725  
 BINAPO [2,2'-Bis(diphenylphosphinyl)-1,1'-binaphthyl] 725  
 BINOL (1,1'-Binaphthol) 283, 725  
 Biocytin [*N*-(+)-Biotinyl-L-lysine] 864  
 Biomass derived organic solvents 97  
 L-erythro-Biopterin 863  
 D-(+)-Biotin 863, 864  
 BIPHEN (3,3'-Di-*tert*-butyl-5,5',6,6'-tetramethyl-1,1'-biphenyl-2,2'-diol) 687, 735  
 B<sub>2</sub>pin<sub>2</sub> [Bis(pinacolato)diboron] 690 729  
 PBN (*N-tert*-Butyl  $\alpha$ -phenylnitron) 294  
 Bis-desmethoxycurcumin 870  
 Bisphenol A (4,4'-Isopropylidene-diphenol) 3, 346, 433  
 Bis(trifluoromethanesulfon)amide (NTf<sub>2</sub>) 730  
 BIS-TRIS [Bis-(2-hydroxyethyl)-amino-tris-(hydroxymethyl)-methane] 118  
 Bitoscanate (1,4-Phenylene diisothiocyanate) 372  
 Blue stone (blue vitriol) (Cupric sulfate) 573  
 BOC-anhydride (Di-*tert*-butyl pyrocarbonate) 139  
 BOC-Cl (Cbz-Cl) 282  
 BOC-ON (2-*t*-Butoxycarbonyloxy-imino-2-phenylacetonitrile) 293  
 Bode-Rovis-Enders Catalysts 714

Boiling point(s), effect of pressure 8, 9, 13, 41, 42  
 of useful gases 59  
 nomogram 40  
 Bombesin (2-L-Glutamin-3-6-L-asparagine-alytesin) 780  
 BMS [Bis(2-mercaptoethyl)sulfone] 118  
 BPD [2-(4-Biphenyl)-5-phenyl-1,3,4-oxadiazole] 412  
 Bradykinin (ArgProProGlyPhe-SerPro-PheArg) 781  
 Brassicasterol [24*R*-(Methylcholesta-5,22*E*-dien-3 $\beta$ -ol)] 843  
 Brilliant Green (4-Dimethylaminotri-phenyl carbinol) 286  
 Bromelain 800  
 Bromocresol Green (3',3'',5',5''-Tetrabromo-*m*-cresolsulfonephthalein) 289  
 Bromocresol Purple (5',3''-Dibromo-*o*-cresolsulfone-phthalein) 289  
 Bromophenol Blue (3,3'',5,5'-Tetrabromophenolsulfonephthalein) 291  
 Bromopyrogallol Red (5,5'-dibromopyrogallolsulfonephthalein) 292  
 Bromothymol Blue (3',3''-Dibromothymolsulfonephthalein) 292  
 Brown & Brown catalyst (Pt/NaBH<sub>4</sub>) 665  
 Brown ring test for NO 576  
 Brucine 865  
 BSA [Bis-(trimethylsilyl)Acetamide] 503  
 BTDA (Benzophenone-3,3',4,4'-tetracarboxylic dianhydride) 277, 937  
 BTFA (Bistrifluoroacetamide) 119  
 BTI [Bis-(trifluoroacetoxy)iodobenzene] 286  
 BTMSA [Bis(trimethylsilyl)acetylene] 503  
 Büchner funnel 14  
 Buchwald ligand 722, 735, 737, 738  
 Buffers, aqueous (Table) 56  
 Bunsen coefficients 36, (Table) 75  
 Butamben (*n*-Butyl *p*-aminobenzoate) 293  
 Butyryne (S-2-Aminobutyric acid) 775

**C**

Cacodylic acid (Dimethylarsinic acid) 505  
 Cactinomycin (Actinomycin C) 858  
 Cadion [1-(4-Nitrophenyl)-3-(4-phenylazophenyl)triazene] 630  
 Calcein [2',7'-Bis-{*N,N*-di(carboxymethyl)aminomethyl}-fluorescein Na salt] 631

- Calcimycin** (Calcium ionophore III) 632  
**Calcium hydride**, in purification 73  
**S-Canavanine** [2-Amino-4-(guanidinoxy)-butyric acid] 781  
**Cannabiscetin** (3,3',4',5,5',7'-Hexahydroxyflavone) 886  
**Capsaicin** (*E-N*-[4-Hydroxy-3-methoxyphenyl)methyl]-8-methyl-6-nonenamide) 295  
**Captopril** (*S*-1-[3-Mercapto-2-methyl-1-oxopropyl]-L-proline) 418  
**(+)-Calarene** (Aristolene) 224  
**Carbon Black** 130  
**Carbostyryl** (2-Hydroxyquinoline) 449  
**Caricaxanthin** (all-*trans*-3*R*- $\beta$ , $\beta$ -Carotene-3-ol) 839  
**Carnitine** ( $\alpha$ -Hydroxy- $\beta$ -*N,N,N*-trimethylaminopropionic acid) 781  
**L-Carnosine** ( $\beta$ -Alanyl-L-histidine) 781  
**Caro's acid** [Oxone] 602  
 $\lambda$ -Carrageenan 819  
**Castro's reagent** (BOP, Benzotriazol-1-yloxytris-(dimethylamino)phosphonium hexafluorophosphate) 502  
**Catapres, Clonidine HCl** [2-(2,6-Dichloroanilino)-2-imidazoline HCl] 867  
**Cathepsin B & D** 801  
**Cation exchange resin(s)** 24, 48, 296, 865  
**CBS catalysts** (Corey-Bakshi-Shibata) 717  
**Cbz-Cl** (Benzylxycarbonyl chloride) 282  
**CDT** [1,1'-Carbonyldi(1,2,4-triazole)] 419  
**CDTA** 230  
**Celite** 7, 15, 21, 22  
**Cellulose**, 15, 20, 21, 22, 25, 26, 44, 48, 48, 49 preparation 25, 26  
**Centrifugation** 15, 765  
**L- $\alpha$ -Cephalin** 892  
**Cephalosporin C** 865  
**Ceruloplasmin** 801  
**CFF** (Cross flow filtration) 923  
**Chalcone** (Benzalacetophenone) 271  
**Calcon carboxylic acid** [3-Hydroxy-4-(2-hydroxy-4-sulfo-1-naphthylazo)naphthalene-2-carboxylic acid] 295  
**CHAPS** (3-[(3-Cholamidopropyl)-dimethylammonio]-1-propane sulfonate betaine) 844  
**CHAPSO** (3-[(3-Cholamidopropyl)-dimethylammonio]-2-hydroxy-1-propane-sulfonate betaine) 844  
**Charcoal**, 21, 23, 48, 79, 82, 83, 86, 87, 104, 112, 131  
 animal 22  
 Darco 22  
 Norit (from Birch wood) 22  
 Nuchar 22  
 ORBO 23  
 purification of 131  
**Chelex-20** 25  
**Chelex-100** 25, 48, 72, 766  
**Chemical hazards** 4-7  
**Chemokines** 801  
**Chiral auxiliaries** 757  
**Chiral chromatography** 49, 50, 60  
**CHIRAPHOS** [2,3-Bis(diphenylphosphino)butane] 725, 726  
**Chirazymes** 801  
**Chloramine-T** (*N*-Chloro-*p*-toluenesulfonamide sodium salt) 506  
**Chloramphenicol** 865, 866  
***p*-Chloranil** (2,3,5,6-Tetrachloro-1,4-benzoquinone) 296  
**Chloranilic acid** (2,5-Dichloro-3,6-dihydroxy-1,4-benzoquinone) 296  
**Chlornaphthazine** 285  
**Chlorogenic acid** [1-(3,4-Dihydroxy-cinnamoyloxy)-D-quinic] acid 298  
**Chlorophenol Red** (3,3'-Dichlorophenolsulfonephthalein) 301  
**Chlorophyll a** 866  
**Chlorophyll b** 866  
**Chloroplatinic acid** 579, 599  
**7-Chlorotetracycline** (Aureomycin HCl) 862  
**Chromatography** 18, 60  
 adsorption 19, 21, 26-28  
 affinity 26, 28, 60, 766, 768  
 automated column 28  
 biospecific 28  
 charge transfer 28  
 chiral 49, 60  
 column 21  
 column for 27  
 covalent 28  
 detectors 27  
 flash 23  
 gas 19, 47  
 GC-MS 29  
 GC-MS-MS 29  
 HPLC 27, 64  
 hydrophobic 28  
 ion exchange 24, 25, 48, 49, 65  
 LC-MS 29  
 LC-MS-MS 29  
 liquid 18  
 mass spectrometry 19, 60, 63, 69  
 metal chelate 28  
 monolithic 28  
 paired-ion (PIC) 23  
 paper 21  
 spot detection in TLC 47  
 stationary phases for GLC 47  
 thin layer (TLC) 20, 69  
 vapour phase 19  
**Chromatotron** 20  
**Chromic acid bath** 3  
**Chromomycin A<sub>3</sub>** 866  
**Chromotropic acid di-Na salt** 512  
**Crotaline** (Monocrotaline) 869  
**Chrysoidine G** (4-Phenylazo-1,3-benzene-diamine monohydro-chloride) 303  
 $\alpha$ -Chymotrypsin 801  
**Cinchonidine** 867  
**Cinchonine** 867  
**Cinchophen** (Atophen) 472  
**Cisplatin** (*cis*-Diammine-dichloro-platinum) 575  
**S-Citrulline** (2-Amino-5-ureidopentanoic acid) 781  
**Citolone** (*N*-Acetyl-*RS*-homocysteine thiolactone) 773  
**Citric acid cycle components** 801  
**Claisen alkali** 572  
**Claisen flask** 10  
**Clavatin** (Patulin) 468  
**Cleaning practices**, 3  
 with chromic acid 3, 5, 6, 15  
 of glassware 3

Cleland's reagent (*RS*-1,4-Dithiothreitol, DTT) 153  
 Clofazimine 423  
 Clostripain 801  
 CTL (Creatol, 2-amino-5-hydroxy-1-methyl-imidazolidin-4-one) 880  
 CNTs (Carbon NanoTubes) 919  
 (–)-Cocaine 867  
 Cocculin (picrotoxin) 255  
 COD (*cis,cis*-Cycloocta-1,5-diene) 234  
 Codeine 867  
 Coenzyme A 867  
 Coenzyme Q<sub>0</sub>, Q<sub>9</sub>, Q<sub>4</sub>, Q<sub>10</sub> 868  
 Colcemide (Demecocine) 868  
 Colchicine {*N*-[(7*S*)-1,2,3,10-Tetramethoxy-5,6,7,9-tetrahydro-9-oxo-benzo[*a*]-heptalen-7-yl]-acetamide} 868  
 Colicin E1 802  
 Collagenase 802  
 Column packings, 10, 11, 23, 27, 28, 49  
 balls or beads 11  
 Dixon rings 11  
 Fenske rings 11  
 helices 11  
 Lessing rings 11  
 Raschig rings 11  
 Wire (Hei-Grid, Heli-Pack) 11  
 Column types  
 Bruun 10  
 Bubble-cap 11  
 Dufton 11  
 Hempel 11  
 Oldershaw 11  
 Podbielniack 11  
 Stedman 11  
 Todd 11  
 Vigreux 11  
 Widmer 11  
 Combinatorial chemistry 94, 95  
 Condensers  
 Allihn 11  
 Bailey-Walker 11  
 coil 12  
 double surface 12  
 Friedrichs 12  
 Graham 12  
 Hopkins 12  
 Liebig 12  
 Othmer 12  
 West 12  
 Wiley 12  
 Congo Red (4B) (cotton red B) 304  
 Congressane (diamantane) 242  
 $\alpha$ -Conhydrine [(±)-2-( $\alpha$ -Hydroxypropyl)piperidine] 447, 448  
 Coniferyl alcohol (4-Hydroxy-3-methoxycinnamyl alcohol) 304  
 Cooling baths 18,  
 (Table) 46  
 COPs catalysts (Cobalt Oxazoline Palladacycles) 680-682  
 Corticotropin 782  
 Coumafene 906  
 Cordite (Erythrityl tetra-nitrate) 822  
 Cordycepin (Adenine 3'-deoxyriboside) 871  
 Cortisol 851  
 Cosbiol, squalane (2,6,10,15,19,23-Hexamethyl-tetracosane) 841  
 Cp (Cyclopentadiene) 234  
 Cp<sup>o</sup> (1,2,3,4,5-Pentamethylcyclopentadiene) 253  
 CPME (Cyclopentyl methyl ether) 241  
 Creatine (*N*-Guanidino-*N*-methylglycine) 782  
 Creatinine (2-Amino-1-methyl-imidazolidin-4-one) 782  
*o*-Cresol Red 305  
*o*-, *m*-, *p*- Cresotic acid (3, 4, & 5-Methylsalicylic acid) 305  
 Crystal Violet Chloride (Gentian violet) 305  
 CT (Computer tomography) 934  
 CTAB (Cetyltrimethylammonium bromide) 131, 753  
 CTFE (Chlorotrifluoroethylene) 134  
 CTL (Creatol; 5-hydroxycreatinine; 2-amino-5-hydroxy-1-imidazolidin-4-one) 880  
 Cupferron (*N*-Nitroso-*N*-phenyl-hydroxylamine ammonium salt) 507  
 Cuproine [Bis(2,9-dimethyl-1,10-phenanthroline)] copper(I) perchlorate) 630

Curcumin [Bis(4-Hydroxy-3-methoxycinnamoyl)methane] 869  
 CVD (Chemical or carbon vapour deposition) 921, 924, 931.  
 Crystallisation 14, 61  
 from the melt 17  
 isothermal 15  
 Cyanocobalamine (Vitamin B<sub>12</sub>) 904  
 Cyclam (1,4,8,11-Tetraazacyclo-tetradecane) 486  
 Cycloleucine (1-Amino-1-cyclopentane carboxylic acid) 776  
 Cyclopentadiene (C<sub>p</sub>, & dimer) 234  
 Cycloserine (2-Amino-3-isoxazolidone) 782  
 Cysteamine (2-Mercaptoethyl-amine) 783  
 Cytarabin (Cytosine-1- $\beta$ -*O*-arabino-furanoside) 871  
 Cytochrome c<sub>1</sub> 802  
 Cytokines 801, 807, 808  
 Cytovene (Gangcyclovir) 878

## D

2,4-D (2,4-Dichlorophenoxyacetic acid) 315  
 DABCO (1,4-Diazabicyclo[2.2.2]-octane) 427  
 DABITC (4-*N,N'*-Dimethylaminoazobenzene-4'-isothiocyanate) 320  
 DADPE (4,4'-Diaminodiphenyl ether) 307, 937  
 -DAG [(–)-2,3:4,6-Di-*O*-isopropylidene-2-keto-L-gulonic acid monohydrate] 822  
 Dactinomycin 859  
 Danishefsky's diene (*trans*-1-Methoxy-3-(trimethylsilyloxy)-1,3-butadiene) 517  
 Danthron (1,8-Dihydroxyanthraquinone) 317  
 Daphnetin (7,8-Dihydroxycoumarin) 431  
 DACH-phenyl Trost ligand 733  
 DACH-pyridyl Trost ligand 734  
 DATD [(+)-*N,N'*-Diallyl tartramide] 137  
 Davy's reagent methyl (2,4-Bis(methylthio)-1,3,2 $\lambda^5$ ,4 $\lambda^5$ -dithiadiphosphetane-2,4-dithione) 502  
 DBED (Dibenzylethylenediamine) 309  
 DBN (1,5-Diazabicyclo[4.3.0]non-5-ene) 427  
 DBP (Dibutyl phthalate) 311  
 DBPC (2,6-Di-*t*-butyl-*p*-cresol) 311  
 DBU (1,8-Diazabicyclo[5.4.0]undec-7-ene) 427  
 DCC (1,3-Dicyclohexylcarbodiimide) 246  
*o,p'*-DDD, Mitotane [1-(2-Chlorophenyl)-1-(4-chlorophenyl)-2,2-dichloroethane] 302

- p,p'*-DDD [2,2-Bis-(*p*-chlorophenyl)-1,1-dichloroethane] 285
- p,p'*-DDE [2,2-Bis-(*p*-chlorophenyl)-1,1-dichloroethylene] 285
- DDQ (2,3-Dichloro-5,6-dicyano-*p*-benzoquinone) 314
- p,p'*-DDT [2,2-Bis-(4-chlorophenyl)-1,1,1-trichloroethane] 285
- DEAD (Diethyl azidodicarboxylate) 95, 142
- Dean-Stark apparatus 77
- DEC, EDCI [1-Diethyl-3-(3-dimethylaminopropyl)-carbodiimide HCl] 95, 148, 915
- Debrox (see UHP) 582
- Decumbin (Brefeldin A) 864
- Dehydrite (Anhydrone) 5, 31, 588
- Deltacortisone (Prednisone) 853
- Demecocine (Colcemide) 868
- Deoxy-L-mannose (L(+)- $\alpha$ -Rhamnose) 831
- Deoxyribonucleic acid (DNA) 766, 767, 768, 770, 863, 872
- DEP (Diethyl pyrocarbonate) 146
- Desiccants 29, 30-32  
 alumina 30  
 aluminium amalgam 30  
 barium oxide 30  
 barium perchlorate 30  
 boric anhydride 31  
 calcium chloride 31  
 calcium hydride 31, 73  
 calcium oxide 31  
 calcium sulfate 31  
 capacity 31  
 common 30  
 copper (II) sulfate 31  
 lithium aluminium hydride 73  
 magnesium amalgam 31  
 magnesium perchlorate 5, 31, 588  
 magnesium sulfate 31  
 molecular sieves 32  
 phosphorus pentoxide 31  
 potassium 31
- potassium carbonate 31  
 silica gel 31  
 sodium 31  
 sodium hydroxide 31  
 sodium-potassium alloy 31  
 sodium sulfate 32  
 sulfuric acid 32
- Dexamethasone (9- $\alpha$ -Fluoro-16- $\alpha$ -methylprednisolone, prednisolone F) 848, 849
- Desmethoxycurcumin 870
- Dextran(s) 818, **821**
- DHFA (7,8-Dihydropteroyl-L-glutamic acid) 872
- Dihydrofolate reductases (assay) 803, 873
- DHP (Dihexadecylphosphate) 509
- Dihydropteridine reductases 803
- DIAD (Diisopropylazodicarboxylate) 147
- Dialysis 765, 766
- Di- and tri-carboxylic acids 872
- Diatomaceous earth 10, 10, 21, 22
- Diazald (*N*-Methyl-*N*-nitroso-*p*-toluenesulfonamide) 352
- 6,6-Dibenzyl-14-crown-4 (lithium ionophore VI, 6,6-dibenzyl-1,4,8,11-tetra-oxa-cyclo-tetradecane) 636
- Dichalcogenides 738
- Dicer (Endonuclease) 623
- Dichloramine-T (*N,N*-Dichloro-*p*-toluenesulfonamide) 312
- Dichloran (2,6-Dichloro-4-nitroaniline) 314
- Dichlorophen [2,2'-Methylenebis(4-chlorophenol)] 314
- Dilauroyl- $\alpha$ -cephalin 873
- Dilituric acid (5-Nitrobarbituric acid) 466
- Dimethyl Yellow (*p*-Dimethyl-aminoazobenzene) 320
- Dimevamide (*dl*-4-Dimethylamino-2,2-diphenylvaleramide) 785
- Dimyristoyl- $\alpha$ -cephalin 874
- DIOP [(2,3-*O*-Isopropylidene)-2,3-dihydroxy-1,4-bis-(diphenyl-phosphino)butane] 742
- Dipalmitoyl- $\alpha$ -L-lecithin 874
- Diphenolic acid [4,4-Bis(4-hydroxyphenyl) valeric acid] 286
- DIPHOS [1,2-Bis-(diphenyl-phosphino)ethane] 726, 752
- Diprophylline [( $\pm$ )-7-(2,3-Dihydroxypropyl)theophylline] 431
- Distearoyl- $\alpha$ -kephalin 874
- Distillation 8, 61  
 azeotropic 13  
 columns 9, 10  
 fractional 10  
 isopiestic or isothermal 14  
 Kügelrohr 14  
 of liquid mixtures 9  
 procedure for 9  
 reflux ratio 10  
 spinning-band 12  
 steam 13  
 techniques of 9  
 types of 10  
 vacuum 12
- Distilling flask 6, 9, 10
- Dithizone (Diphenylthiocarbazone) 331
- Diuron [3-(3,4-Dichlorophenyl)-1,1-dimethyl urea] 315
- Djenkolic acid (*S,S'*-Methylene-bis-L-cysteine) 785
- DLS (Dynamic Light Scattering) 908, 937
- DMAB (Ehrlich's reagent) (*p*-Dimethylaminobenzaldehyde) 320
- DMAC (*N,N*-Dimethylacetamide) 148
- DMAP (4-Dimethylaminopyridine) 220, 413, **718**
- DMBA (7,12-Dimethylbenz[*a*]-anthracene) 322
- DMDT "Methoxychlor" (1,1-Bis[*p*-methoxyphenyl]-2,2,2-trichloroethane) 347
- DME (Dimethoxyethane, glyme) 147
- DMEU (DMI; *N,N'*-Dimethylethylene urea) 433, 873
- DMEDA (*N,N'*-Dimethylethylene-diamine) 718
- DMF (*N,N*-Dimethylformamide) 150
- DMI (DMEU; 1,3-Dimethyl-2-imidazolinone) 433
- DMP (Dess-Martin Periodinane) 494, 495
- DMPSCI (Dimethylphenylsilyl chloride) 510
- DMSO (Dimethylsulfoxide) 152
- DMT (4,4'-Dimethoxytrityl chloride) 320
- DNA 803  
 from plasmids 803

purification 770, 767  
 DNPH (2,4-Dinitrophenylhydrazine) 23, 95, 252, 327, 896  
 DNPH-Celite 130, 133, 272  
 DNPH-cartridges 23, 95  
 Dolichol 875  
 Domoic acid [4-(2-Carboxyhexa-3,5-dienyl)-3-carboxymethylproline] 875  
 DOPA, (EUODOPA, LEVO-DOPA, 3,4-Dihydroxyphenylalanine) 784, 785  
 2,4-DP [ $\alpha$ -(2,4-Dichlorophenoxy)propionic acid, Dichloroprop] 315  
 DPEN (1,2-Diamino-1,2-diphenylethane) 739, 740  
 DPPA (Diphenylphosphorazidate) 511  
 Dppb [1,4-Bis-(diphenylphosphino)butane] 726  
 Drierite (Calcium sulfate) 21, 31, 567  
 Drying, 22, 29-32, 62  
   liquids for drying  
   pistols 50  
   removal of water 29  
   of solvents 29, 52  
 Drying agents for organic compounds 52  
 DTBCL<sub>2</sub> (Di-*t*-butyl dichlorosilane) 508  
 Dtbpy (4,4'-Di-*tert*-butyl-2,2'-bipyridyl) 734  
 DTE (Dithioerythritol) 153  
 DTPA (DEPTAPAC) (Diethylenetriaminepentaacetic acid) 143  
 DTT (Dithiothreitol) 153  
 Dumasine (Cyclopentanone) 240  
 DWNTs (Double Walled Carbon NanoTubes) 924  
 Dyphylline (Diprophylline) 431

## E

EA (Electric arc) 925  
 Ebullioscopic Constant 9  
 $\alpha$ - and  $\beta$ -Ecdyson 849  
 EDCI [1-Ethyl-3-(3-dimethylaminopropyl)carbodiimide HCl] 95, 148, 915

EDDA (Ethylenediaminediacetic acid) 161  
 EDTA (Ethylenediaminetetraacetic acid) 161  
 EDTA H<sub>2</sub>Na<sub>2</sub> salt 636  
 EDTA Na<sub>2</sub>Ca salt 636  
 EDTA Na<sub>2</sub> Mg salt 637  
 EDX (Energy Dispersive X-ray Spectroscopy) 909, 921, 936  
 EEDQ (2-Ethoxy-1-ethoxycarbonyl-1,2-dihydroquinoline) 438  
 EGTA [Ethylene glycol bis( $\beta$ -aminoethylether)-*N,N'*-tetraacetic acid] 161  
 Ehrlich's Reagent 320  
 Electrophoresis 29, 62, 766-769  
 Enzymes  
   [EC 1.1.1.5] 798  
   [EC 1.1.1.8] 806  
   [EC 1.1.1.27] 808  
   [EC 1.1.1.30] 807  
   [EC 1.1.1.49] 805  
   [EC 1.1.3.4] 805  
   [EC 1.2.1.12] 806  
   [EC 1.5.1.3] 803  
   [EC 1.6.99.7] 803  
   [EC 1.9.3.1] 803  
   [EC 1.10.2.2] 815  
   [EC 1.14.17.1] 804  
   [EC 1.15.1.1] 802  
   [EC 2.4.1.11] 806  
   [EC 2.4.2.30] 799  
   [EC 2.5.1.18] 806  
   [EC 2.7.1.30] 806  
   [EC 2.7.1.40] 812  
   [EC 2.7.1.70] 811  
   [EC 2.7.6.1] 811  
   [EC 2.7.7.9] 815  
   [EC 2.7.7.49] 813  
   [EC 3.1.1.34] 808  
   [EC 3.1.27.5] 813  
   [EC 3.1.3.1] 810  
   [EC 3.1.3.5] 810  
   [EC 3.2.1.8] 815  
   [EC 3.2.1.17] 809  
   [EC 3.2.1.23] 805  
   [EC 3.2.1.28] 815  
   [EC 3.2.1.35] 807  
   [EC 3.2.1.50] 798  
   [EC 3.2.1.52] 798  
   [EC 3.4.4.13] 814  
   [EC 3.4.13.11] 812  
   [EC 3.4.14.5] 803

  [EC 3.4.17.1] 800  
   [EC 3.4.21.1] 801  
   [EC 3.4.21.31] 815  
   [EC 3.4.21.62] 814  
   [EC 3.4.22.1] 801  
   [EC 3.4.22.2] 810  
   [EC 3.4.22.8] 801  
   [EC 3.4.23.1] 810  
   [EC 3.4.23.5] 801  
   [EC 3.4.24.3] 802  
   [EC 3.4.33.4] 800  
   [EC 3.5.2.2] 807  
   [EC 3.5.4.4] 896  
   [EC 3.14.15.1] 800  
   [EC 4.1.3.3] 798  
   [EC 4.2.1.1] 800  
   [EC 6.2.1.1] 798  
   [EC 6.2.1.2] 798  
   [EC 6.5.1.3] 815  
 Ellagic acid (4,4',5,5',6,6'-Hexahydroxydiphenic acid dilactone) 332  
 EM (Electron microscopy) 920  
 Emodine (see Archin) 332  
 Ephedrine (2-Methylamino-1-phenylpropanol) 875  
 Epinephrine (L-Adrenaline) 860  
 Epsom salt (Magnesium sulfate) 588  
 Ergosterol (Provitamin D<sub>2</sub>) 849  
 $\alpha$ - and  $\beta$ -Erythroidine 875, 876  
 Erythroidine  $\alpha$ ,  $\beta$ , & HCl's 875-876  
 Esculetin, cichorigenin (6,7-Dihydroxycoumarin) 438  
 Estrone (1,3,5-Estratrien-3-ol-17-one, Folliculin) 850  
 Ethacridine (6,9-Diamino-2-ethoxyacridine) 426  
 Ethanol 156  
 Ethidium bromide 53, 876  
 Ethionamide 439  
 Ethotoin [( $\pm$ )-3-Ethyl-5-phenylhydantoin] 439  
 Etiocholane (5 $\beta$ -Androstane) 842  
 Ethoxyquin (1,2-Dihydro-6-ethoxy-2,2,4-trimethylquinoline) 438  
 Ethyl Red [2-(4-Diethylaminophenylazo)benzoic acid] 334  
 Eu(dcm)<sub>3</sub> 660  
 Euf flavin (Acriflavin Mixture) 400  
 Eugenol (4-Allyl-2-methoxyphenol) 334  
 Eu(tfc)<sub>3</sub> 660  
 Eutonyl (Pargyline HCl) 366  
 Exaltone (Cyclopentadecanone) 234  
 Exonucleases 804

## F

FA folic acid (Pteroyl-*S*-glutamic acid) 877

facac (tfacac; 1,1,1-Trifluoro-2,4-pentanedione) 214, 937

FAD (Flavin Adenine Dinucleotide) 877

FEM (FluorescePAe Emission Microscopy) 909, 919, 920

Fenazox (Azoxybenzene) 271

Ferritin 804

Fibrinogen 804

Fibronectin 805

Fieser's solution 594

Filters, micro 44  
Whatman 43

Filtration 15

Flavanol (3-Hydroxyflavone) 446

Florisil 15

Flow chemistry 93

Fluoranthene (Benzo[*j,k*]fluorene) 334

Fluorescein [9-(*o*-Carboxyphenyl)-6-hydroxy-3*H*-xanthene-3-one] 53, 440

Fluoresceinamine (Mixture of 5- and 6-aminofluorescein) 53, 440

Fluorescein isothiocyanate isomer I (5-Isothiocyanato isomer) 53, 440

Fluorous chemistry 97, 102

Fluram (Fluorescamine, 4-phenyl-spiro-[furan-2(3*H*)-1-phthalan]-3,3'-dione) 53, 441

FMN (Flaniv adenine dinucleotide) 877

Fmoc (9-Fluorenylmethylcarbonyl-) 94

Fmoc-Cl (9-Fluorenylmethyl chloroformate) 335

Folliculin (1,3,5-Estratrien-3-ol-17-one) 850

Follitropin (Follicle stimulating hormone) 805

Forskolin, Colforsin 441

Forthane (2-Amino-4-methylhexane) 115

F-propene (Perfluoropropene) 145, 169

Fractogels 26, 49

Freeze-pump-thaw and purging 37

Fremy's Salt (Potassium nitroso sulfonate) 602

Freon 12 (Dichlorodifluoro-methane) 140

Freon 13B1 (Bromotrifluoro-methane) 121

FSH (Follicle Stimulating Hormone) 805

5FU (5-Fluorouracil) 440

Fullerene C<sub>60</sub>- $\gamma$ -lactone esters 915

Fullerene[60]/PC [Poly(bis-phenol A)carbonate] 915

Fullerenes 909-913

Fumagillin 441

Fumarine 475

Functionalised CNTs

Functionalised fullerenes 913-919

Fungicidin (Nystatin) 890

Fusaric acid (5-*n*-Butylpyridine-2-carboxylic acid) 443

Fuschin, rosaniline HCl 379

Fuming nitric acid 594

## G

GA (Gum Arabic) 934, 937

GABA (*S*-2-Aminobutyric acid) 776

GABOB [*R* (L-)-4-Amino-3-hydroxybutyric acid] 777

Galactitol 822

Galvinoxyl [2,6-Di-*tert*-butyl- $\alpha$ -(3,5-di-*tert*-butyl-4-oxo-2,5-cyclohexadiene-1-ylidene)-*p*-tolylxy] 336

Gangcyclovir 878

Gas chromatography 19, 47, 63

GDH (L-Glycerol-3-phosphate dehydrogenase) 806

Gel electrophoresis 768

Gel filtration 26, 63, 766

Gentian violet 305

Germanium tetroxide [14165-55-0] 638-639

GIXRR (Grazing Incident X-Ray Refractometry) 909

Girard Reagent T (*N,N,N*-Trimethylglycinehydrazide chloride) 797

Glass powder 578

Globulin G<sub>1</sub> 809

Glovebox techniques 37

D-Glucitol 824

Glycamine 824

Glycocyanine (*N*-Guanylglycine) 787

GA-AuNP (Au nano-particles stabilised by gum Arabic) 934

Graded adsorbents 21, 48

Gramicidin A, C, & S 788

Gramine (3-Dimethylamino-ethylindole) 443

Graphane 930

Graphene 929

Green chemistry 93, 97, 98

Green vitriol (Ferrous sulfate) 576

Griseofulvin 443

Grubbs catalyst 705, 706

Guaiaic acid [4,4'-(2,3-Dimethyl-1-butene-(1,4-diyl)-bis-(2-methoxyphenol)] 337

Guaiazulone (1,4-Dimethyl-7-isopropylazulene) 337

Guanabenz (1-[2,6-Dichlorobenzylideneamino]guanidine WY-8677) 337

Guanazole (3,5-Diamino-1,2,4-triazole) 427

Gypsum (Calcium sulfate) 567

## H

Hematin 880

Haemoglobin A 806

Haematoporphyrin IX 880

Hagemann's ester (4-Carboethoxy-3-methyl-2-cyclohexen-1-one) 226

Harmaline (7-Methoxy-1-methyl-4,9-dihydro-3*H*- $\beta$ -carboline) 443

Harmane (2-Methyl- $\beta$ -carboline) 443

Harmine (7-Methoxy-1-methyl-9*H*-pyrido[3,4-*b*]indole) 444

Hazards,  
Chemical 4-7  
Nano materials 939-94

HBTU (*O*-Benzotriazol-1-yl-*N,N,N,N'*-tetramethyluronium F<sub>6</sub>PO<sub>4</sub><sup>-</sup>) 410

Heating baths 43

HEDTA [*N*-(2-Hydroxyethyl)ethylene diaminetriacetic Acid] 172

Heimgartner's reagent [2,4-Bis(*p*-tolylthio)-1,3,2 $\lambda^5$ ,4 $\lambda^5$ -dithiadiphosphetane-2,4-dithio] 503

Hemimellitic acid (Benzene-1,2,3-tri-carboxylic) 274

Hesperetin (3',5,7-Trihydroxy-4'-methoxy-flavanone) 444

Hesperidin (Hesperetin 7-rhamnoside) 444

Heteroauxin (Indole-3-acetic acid) 451

HEPES [Na piperazine-*N,N'*-bis(2-ethane-sulfonate)] 72, 446

HETP, height equivalent to a theoretical plate 9

Hexesterol [3,4-Bis-(4-hydroxyphenyl)-hexane] 285

hfacac (1,1,1,5,5,5-Hexafluoro-2,4-pentanedione) 169

HFIP (1,1,1,3,3,3-Hexafluoropropan-2-ol) 170

HIMDA (*N*-Hydroxyiminodiacetic Acid) 172

HiPco (High Pressure Carbon Monoxide Disproportionation)

Hippuric acid (Benzoyl glycine) 780

Histones 807

HMDB [Hexamethyl(Dewar)-benzene] 249, 337-338

HMDS (Hexamethyldisilazane) 514

HMDSO (Hexamethyldisiloxane) 514

HMH (5-Hydroxy-1-methylimidazol-idine-2,4-dione) 881

HMPA (Hexamethylphosphoric triamide) 133, 433, 514

HMPT Hexamethylphosphorous triamide) 514

HOBt (1-Hydroxybenzotriazole) 95, 446, 915

HPLC, high performance liquid chromatography 27, 60, 64, 579  
gel filtration 26, 63  
reverse-phase 27

HPTLC (High performance thin layer chromatography) 20

5-HT [Serotonin HCl; 3-(2-aminoethyl)-5-hydroxyindole HCl] 897

Hünig's Base (Diisopropylethylamine) 147

Hyamine (1622) 275

Hydralazine (1-Hydrazino-phthalazine) 445

Hydranal 2, 30

Hydantoin (2,4-Dihydroxy-imidazole) 873

Hydrogen ionophore II 467

Hydrogen ionophore IV 467

Hydroxyamphetamine [4-(2-Aminopropyl)phenol] 267

Hydroxylapatite 26, 766

Hyflo-supercel 15

Hypo (Sodium thiosulfate) 616

Hypericin (4,5,7,4',5',7'-Hexahydroxy-2',2'-dimethylnaphthodianthrone) 342

Hyperol (see UHP) 582

## I

Ibotenic acid [ $\alpha$ -(3-Hydroxy-5-isoxazolyl)glycine] 790

Ibuprofen (4-*iso*Butyl- $\alpha$ -methylphenylacetic acid) 342

IBX [1-Hydroxy-1,2-benziodoxol-3(1*H*)-one; 2-iodoxybenzoic acid] 494

ICC (Iterative cross-coupling) 745

Impurities, sources of 3, 4

Ionic equilibria 34, 64

Ion-exchange 25, 65, 72  
bead form 24, 26, 49  
Cellex CM 26  
Cellex D 26  
celluloses 25, 26, 48  
Chelex 25, 48  
resins 24, 25, 43  
Sephadex 25, 26  
Sephacrose 26

Ionic liquids (ILs) 96, 97, 101, 755

Ionisation constants 8, 24, 34, 103, 555, 771  
pK and temperature 35  
pK and solvent 36

IPNBSH (*N*-Isopropylidene-*N'*-2-nitrobenzenesulfonyl hydrazide) 345

Interferons 807

Interleukins 807-808

Ionic liquids 96, 755

Iproniazid (1-Isonicotinyl-2-isopropylhydrazide) 453

Ishikawa's Reagent (PPDA) 145

Isoelectric focusing 766-767

Isoniazide (Isonicotinic acid hydrazide) 453

Isosarsasapogenin 854

isothermal distillation 14

Isoxanthopterin (2-Amino-4,7-dihydroxypteridine) 453

## J

Jacobsen's catalysts 675

Jacobsen's ligand 724

Janus Green B (3-Dimethylamino-[4-dimethylaminoazo]-5-phenylphenazonium chloride) 453

Janus Red B 345

Juglone (5-Hydroxy-1,4-naphthoquinone) 341

Julolidine (2,3,6,7-Tetrahydro-1*H*,5*H*-benzo[*ij*]quinolizidine) 454

## K

Kanamycin B (Beknamycin) 882

Kemp's acid (Trimethylcyclo-hexane-1,3,5-tricarboxylic acid) 256

Kerosene 177

Ketone Moschus (4-*tert*-Butyl-2,6-dimethyl-3,5-dinitroacetophenone) 345

Kinetin (6-Furfurylaminopurine) 442

Koser's reagent (Hydroxy(tosyloxy)-iodobenzene) 342

Koshland's reagent (2-Hydroxy-5-nitrobenzyl bromide) 341

KPEP (K Phosphoenolpyruvate) 892

## L

Laboratory techniques 65

Laburnine 424

Lactaminic acid (NANA) 857

Lactide (3,6-Dimethyl-1,4-dioxan-2,5-dione) 432

Lactoferrin 808

Lactoflavin (Vitamin B<sub>2</sub>) 904

LAH (Lithium aluminium hydride) 73, 586

Lanthanide (Europium) shift reagents 177, 660

- Laughing gas (Nitrous oxide) 594  
 Lauth's Violet 490  
 Lawesson's reagent (2,4-Bis(4-methoxyphenyl)1,3,2,4-dithia-diphosphetane-2,4-disulfide) 515  
 Lazier catalyst (Copper chromite) 663  
 Lectins 798, 808  
 Levallorphan [1-*N*-Allyl-3-hydroxy-morphinan] 402  
 Levocarnitine 781  
 Levomethadone 883  
 Lignoceric acid (Tetracosanoic acid) 204  
 Ligroin 17  
 Lipoproteins 808  
 Lipowitz alloy 623  
 Liquid crystal catalysts (Imidazolium salts) 755  
 Lithium aluminium hydride, in purification 73  
 Lithium hydride 587  
 D-Luciferin (from firefly) 883  
 Lumichrome (7,8-Dimethylalloxazine) 455  
 Lumiflavin (7,8,10-Trimethylbenzo[*g*]pteridine-2,4-(3*H*,10*H*)-dione) 883  
 Luminol (5-Aminophthalazin-1,4-dione) 455  
*dl*-Lupinine (1-Hydroxymethyl-octahydroquinolizine) 455  
 Lutein (Xanthophyll) 841  
 (+)-Lysergic acid 456  
 Lysozyme 809
- ## M
- MAAN (*N*-Methyleneamino-acetonitrile) 184  
 Macleyine (Protopine) 475  
 McLeod type gauge 12  
 Mafenide HCl (4-Aminomethylbenzenesulfonamide hydrochloride) 265  
 Magenta I 379  
 Magneson -I 359  
 Magon, Xylidyl Blue 642  
 Malachite green (carbinol) 346  
 Maltol (3-Hydroxy-2-methyl-Mullein-4-pyrone) 456  
 Maneb [Mn ethylenebis(dithiocarbamate)] 516  
 Marsilide (1-Isonicotinyl-2-isopropylhydrazide phosphate) 453  
 Martius Yellow (2,4-Dinitro-1-naphthol) 326  
 MCPA (4-Chloro-2-methylphenoxy acetic acid) 299  
 MCPBA (*m*-Chloroperbenzoic acid) 301  
 MDEA (*N*-Methyldiethanolamine) 184  
 Medrol (6- $\alpha$ -Methylprednisolone) 852  
 Medroxyprogesterone (17 $\alpha$ -Hydroxy-6 $\alpha$ -methylprogesterone) 852  
 Mega-9 (*N*-Nonanoyl-*n*-methylglucamine) 830  
 Mega-10 (*N*-Decanoyl-*N*-methylglucamine) 819  
 Meglumine (*N*-Methyl-D(-)-glucamine) 829  
 Meglutol (3-Hydroxy-3-methylglutaric acid) 172  
 MEK (Methyl ethyl ketone) 123  
 Meldrum's acid (2,2-Dimethyl-1,3-dioxane-4,6-dione) 493, 432  
 Melphalan (4-[Bis-(2-Chloroethyl)-amino]-L-phenylalanine) 792  
 MEMS 935  
 MEMCl (2-Methoxyethoxymethyl-chloride) 181  
 Meprobamate [2,2-Di(carbamoyloxymethyl)pentane] 179  
 Mercurochrome [2',7'-Dibromo-4-(hydroxymercurio)fluorescein di-Na salt] 643  
 MES [2-(*N*-Morpholino)ethanesulfonic acid] 465  
 MESNA [Sodium 2-mercaptoethanesulfonate] 531  
 Mestanolone (17 $\beta$ -Hydroxy-17 $\alpha$ -methyl-3-androsterone) 852  
 Mesterone (17 $\alpha$ -Methyltestosterone) 853  
 Metacetamol (3-Acetamidophenol) 258  
 Metabolonics 769  
 Metal hydrides 73, 89  
 Metal impurities, by complexation 73  
 by ion-exchange resins 72  
 by precipitation 72  
 removal by extraction 73  
 removal from reagents 73  
 removal of lead 72  
 removal of iron 72  
 Metalphthalein (*o*-Cresolphthalein complexon) 304  
 Metals, trace analysis 90  
*R*(-)-Methadone (*levo*-6-Dimethylamino-4,4-diphenylheptan-3-one) 883  
 ( $\pm$ )-Methadone HCl 884  
 Methotrexate, amethopterin 860  
 Methoxalen (Xanthotoxin) 906  
 5-Methoxypsoralen (Bergapten) 863  
 9-Methoxypsoralen 906  
 Methylarenes 349  
 Methyl dopa [3-(3,4-Dihydroxyphenyl)-2-methyl-L-alanine] 785  
 Methylene Blue [3,7-Bis-(dimethylamino)phenothiazin-5-ium chloride] 458  
 Methylene Green [3,7-Bis-(dimethylamino)-4-nitrophenothiazin-5-ium chloride] 458  
 Methyl Meldrum's acid (2,2,5-Dimethyl-1,3-dioxane-4,6-dione) 493  
 Methyl Orange (Sodium 4,4'-dimethylaminophenylazobenzene-sulfonate) 517  
 Methyl Red (4-Dimethylaminoazobenzene-2'-carboxylic acid) 651  
 5-Methyltetrahydrofolic acid (prefolic acid) 885  
 Metol (*p*-Methylaminophenol sulfate) 348  
 Metrazol, Cardiazol (3a,4,5,6,7,8-Hexahydro-1,2,3,3a-tetraazaazulene) 464  
 Mevastatin (Compactin) 868  
 MF (Magnetic Fluid) 934  
 MH (1-Methylhydantoin; 1-methylimidazolidin-2,4-dione) 884  
 MIB (Nugent's reagent) [(2*S*)-3-exo-(Morpholino)isborneol] 695  
 Michler's ketone [4,4'-Bis(dimethylamino)benzophenone] 353  
 Michler's base 384  
 Microwave technologies 93  
 MIDA (*N*-Methylimino-diacetic acid) 731  
 MIDA boronic esters (Formation & hydrolysis) 745  
 MIDA-disodium salt 745  
 Mithramycin A 885

Mixed solvents **17, 21, 39, 45**  
 MMPP (Monoperphthalic acid mono Mg salt) **643**  
 Mohr's salt (Ammonium ferrous sulfate) **558**  
 Molecular sieves **32, 67**  
 MOMCl (Chloromethyl methyl ether) **133**  
 Monacrin (9-Aminoacridine HCl) **402**  
 Monocrotaline (Crotaline) **869**  
 Monolithic chromatography **28**  
 MONO-TRIS [2-Hydroxyethylimino-tris(hydroxymethyl)methane] **172**  
 Monuron [3-(4-Chlorophenyl)-1,1-dimethylurea] **302**  
 Morin (2',3,4',5,7-Pentahydroxyflavone) **464**  
 (-)-Morphine **464**  
 Mosher's acid ( $\alpha$ -Methoxy- $\alpha$ -trifluoromethylphenylacetic acid) **348**  
 MPD (*R,S*-2-Methyl-2,4-propanediol) **186, 750**  
 MPTP (1-Methyl-4-phenyl-1,2,3,6-tetrahydropyridine hydrochloride) **460**  
 MRFM (Magnetic Resonance Force Microscopy) **909**  
 MRI (Magnetic Resonance Imaging) **908**  
 $\alpha$ -MSH ( $\alpha$ -Melanotropin) **791**  
 $\beta$ -MSH ( $\beta$ -Melanotropin) **792**  
 MTAD (4-Methyl-1,2,4-triazoline-3,5-dione) **463**  
 MTB (Methyl thymol blue, Na) **518**  
 MTBE (Methyl *t*-butyl ether) **128**  
 MTO (Methyltrioxorhenium) **694**  
 MTPA ( $\alpha$ -Methoxy- $\alpha$ -trifluoromethylphenylacetic acid) **348**  
 MTT (Methyl thiazolyltetrazolium bromide blue) **489**  
 Muramic acid [*R*-2(2-Amino-2-deoxy-D-glucose-3-yloxy)-propionic acid] **885**  
 Muramidase (Lysozyme) **809**  
 Muriatic acid (conc HCl) **581, 582**  
 Muscimol (5-Aminoethyl-3[2*h*]-

isoxazolone) **886**  
 MWCNTs (MWNTs) (Multi Walled Carbon Nano Tubes) **920**  
 Mycobacinin  $\epsilon$ -[2-(4-Thiazolidinone)]hexanoic acid) **900**  
 Mycostatin (Nystatin) **890**  
 Myoglobin **810**  
 Myricetin (Cannabiscetin) **866**  
 Myronate K (Sinigrin) **832**

## N

Nabam [Na ethylenebis(dithiocarbamate)] **512**  
 NAD (DPN) ( $\beta$ -Nicotinamide adenine dinucleotide) **887**  
 NADH, reduced NAD **888**  
 NADP (TPN) ( $\beta$ -Nicotinamide adenine dinucleotide phosphate) **888**  
 NADPH, reduced NADP **888**  
 Nájera catalysts (Oxime palladacycles) **697**  
 Naloxone HCl **886**  
 Naltrexone HCl **887**  
 NAM (*N*-(9-Acridinyl)maleimide) **400**  
 NAMA (*N*-Acetylmuramic acid) **886**  
 NANA (*N*-Acetylneuraminic acid) **857**  
 NaPAA (Sodium polyacrylate) **653**  
 Naptalam **357**  
 Naphthazarin (5,8-Dihydroxy-1,4-naphthoquinone) **354**  
 Narcan (Naloxone HCl) **886**  
 Naringin (4',5,7-Trihydroxyflavone 7-rhamnoglucoside) **829**  
 NBD (2,5-Norbornadiene) **253, 745**  
 NBD-chloride (7-Chloro-4-nitrobenzoxadiazole) **421**  
 NBSH (*o*-Nitrobenzenesulfonyl hydrazide) **359**  
 NC-AFM (Non Contact-Atomic Force Microscopy) **908**  
 ND (Nanodiamond) **932**  
 NDELA (*N*-Nitrosodiethanolamine) **190**  
 Nebularin (Purine-9- $\beta$ -ribofuranoside) **895**  
 Nembutal (5-Ethyl-5-1'-methylbutyl barbituric acid, Na salt) **469**

Neocuproine (2,9-Dimethyl-1,10-phenanthroline) **433**  
 Neo-oxine (4-Hydroxyacridine) **445**  
 Neophyl chloride (2-Chloromethyl-2-phenylpropane) **299**  
 D(+)-Neopterin **887**  
*cis*-, *trans*-, *cis/trans*-Nerolidol **189**  
 Norephedrine [2-Amino-1-phenyl-1-propanol] **889**  
 NF (Nanofiltration) **767, 933**  
 Niacin (Nicotinamide) **465, 466**  
 Nile Blue A (a Benzophenoxazinium sulfate dye) **466**  
 Nigrin (Streptonigrin) **898**  
 Nioxime (Cyclohexane-1,2-dione dioxime) **231**  
 NIR (Near Infra Red) **909, 920, 922, 925, 927**  
 Nisin (Polypeptide) **793**  
 Nitric acid (anhydrous) **593**  
 Nitrin (2-Aminobenzaldehyde phenylhydrazone) **262**  
 Nitromethane **190**  
 Nitrososulfuric acid **594**  
 NMN ( $\beta$ -Nicotinamide mononucleotide) **888**  
 Nonactin (Ammonium ionophore I) **627**  
 Norpseudoephedrine **889**  
 Norspermidine (Bis-[3-amino-propyl]-amine) **137**  
 NPS-Cl (2-Nitrobenzenesulfonyl chloride) **359**  
 NTA [Nitrilotriacetic acid; tris-(carboxymethyl)amine] **190**  
 DTBNpP (Di-*tert*-butylneopentylphosphine) **735**  
 NTF<sub>2</sub> [Bis-(trifluoromethane-sulfon)amide] **708, 730, 755,**  
 Nucleotide thiophosphates **890**  
 Nylon powder **191**  
 Nystatin (Mycostatin, Fungicidin) **890**

## O

Octyl- $\beta$ -D-glucopyranoside **830**  
 ODA (Octadecylamine) **191, 926**  
 ODCB (*o*-Dichlorobenzene) **312, 912**  
 Oil of vitriol, H<sub>2</sub>SO<sub>4</sub> **618**  
 Oleum (Fuming H<sub>2</sub>SO<sub>4</sub>) **619**  
 OP catalysts (Oxime Palladacycle dimers) **697**  
 Optochin (Ethyl hydrocupreine hydrochloride) **439**  
 ORBO cartridges **23**  
 Orcine **365**  
 Organocatalysis **92, 712**  
 Ostwald coefficient **36**  
 (Table) **58**

- Ouabain 853  
 Overman Rearrangement 672, **680**  
 Oxacillin (5-Methyl-3-phenyl-4-isoxazolyl-penicillin sodium salt) 890  
 Oxine (8-Hydroxyquinoline) 449  
 Oximes (Preparation) 697  
 Oxiracetam 890  
 Oxone (Caro's acid, Potassium peroxymonosulfate) 602  
 Oxycinchophen (3-Hydroxy-2-phenyl-cinchonic acid) 447  
 Oxytetracycline (Terramycin) 898  
 Oxytocin 793
- ## P
- PABS (Poly-*p*-aminobenzene-*m*-sulfonic acid) 926  
 PAGE (Polyacrylamide gel electrophoresis) 768, 769  
 L- $\beta$ -Palmitin 891  
 PAN [4-(2-Pyridylazo)naphthol] 479  
 Pantothenic (Na and Ca Salt) 891  
 Paper chromatography 21  
 PAR [4-(2-Pyridylazo)resorcinol] 480  
 Paracetamol (4-Acetamidophenol) 258  
 Praxadine (1-Amidinopyrazole hydrochloride) 477  
 Paraffin wax 194  
 Paraquat 2HCl (1,1'-Dimethyl-4,4'-bi-pyridylum dichloride) 432  
 Patulin, Clavatin [4-Hydroxy-4*H*-furo(3,2-*c*)pyran-2(6*H*)-one] 468  
 PBN (*N*-*t*-Butyl  $\alpha$ -phenyl nitron; free radical trap) 294  
 Pectic acid 830  
 Pectin 830  
 PEG (Polyethylene glycol) 893-894  
 Penicillamine 891  
 Penicillin G 862-863  
 Pentadecanolide (1-oxacyclohexadecan-2-one) 194  
 Pentalin (Pentachloroethane) 194  
 Pentaphene (2,3,6,7-Dibenzphenanthrene) 309  
 Pentobarbital (5-Ethyl-5-1'-methylbutyl barbituric acid) 469  
 Perex test for peroxides 84  
 Pericyazine {10-[3-(4-Hydroxy-1-piperidinyl)-propyl]-10*H*-phenothiazine-2-carbonitrile} 470  
 Petroleum ether 17, 39, 45, **196**  
 Petroleum spirit 17  
 PFBHA HCl [*O*-(2,3,4,5,6-Pentafluorobenzyl)hydroxyl-amine hydrochloride] 366  
 Phenacetin (*p*-Methoxyacetanilide) 259  
 Phenamisol (Phenyl 4-amino-salicylate) 370  
 Phenozin (Zn phenol-*o*-aulfonate) 554  
 Phlorizin 831  
 Phloron (2,6-Dimethyl-1,4-benzoquinone) 322  
 Phosferrox ligands 699, 700  
 Phospholipids 893  
 Phosphoproteins 811  
 Phosphorus sesquisulfide P<sub>4</sub>S<sub>3</sub> 598  
 Photol (Metol) (*p*-Methylamino-phenol sulfate) 348  
 Physostigmine, Eserine 438  
 PIC, chromatography 23  
 Picrotoxin (cocculin) 255  
 Pinacolborane (4,4,5,5-Tetramethyl-1,3,2-dioxaborolane) 747  
 PIPES (Piperazine-*N,N'*-bis(2-ethanesulfonic acid) 474  
 PIPES-Na salt 532  
 Piracetamide (2-Oxo-1-pyrrolidine acetamide) 893  
 Pituitary Growth Factor 811  
 Pixyl chloride (9-Chloro-9-phenylxanthene) 421  
 pK 34, 103, 555, 771  
 pKa and pKb 35 and temperature 35 and solvent 36  
 Plasmids (DNA) 803, 811  
 Platinum (II) chloride 599  
 Platinum Sponge 579, 598  
 Plicamycin (Mithramycin A) 885  
 PLP, codocarboxylase 896  
 Plumbagin (5-Hydroxy-2-methyl-1,4-naphthaquinone) 340  
 PMDBD (3,3,6,9,9-Pentamethyl-2,10-diazabicyclo[4.4.0]dec-1-ene) 469  
 PMSF (Phenyl methanesulfonyl fluoride) 374  
 PNBP [4-(4-Nitrobenzyl)pyridine] 466  
 P<sub>3</sub>NO [4-(3-Phenylpropyl)pyridine 1-oxide] (cf 471), 675, **747**  
 Polybrene (1,5-Dimethyl-1,5-diazaundecamethylene polymethobromide) 149  
 Polyethylene glycol (PEG) 893  
 Polymer supported reactants 95  
 Polypeptides, purification 768  
 Polyprenols (Dolichols) 875  
 Polystyrene 377  
 Potassium borohydride, in purification 74  
 Potassium chromate 600  
 Potassium hexachloroiridate (III) (K<sub>3</sub>IrCl<sub>6</sub>) 601  
 Potassium nickel sulfate hexahydrate 602  
 Potassium nitrate (saltpetre) 602  
 Potassium thiosulfate hydrate (75% aqueous solution)] 604  
 PPDA (Perfluoropropyl-diethylamine) 145  
 PPD (2,5-Diphenyl-1,3,4-oxadiazole) 435  
 PPEI-EI [Poly(propionylethylenimine-co-ethylenimine)] 928  
 P-Phos [2,2',6,6'-Tetramethoxy-4,4'-bis(diphenylphosphino)-3,3'-bipyridine] 749  
 PPO (2,5-Diphenyloxazole) 435  
 PQQ (Methoxantin coenzyme) 884  
 Prazosin HCl 894  
 Prefixes for quantities 59  
 Prefolic A 885  
 Preparative GC 19, 63  
 Primaquine diphosphate 475  
 Pristane, norphytane (2,6,10,14-Tetramethylpentadecane) 208  
 Procaine HCl (Novocain, 2-diethylaminoethyl-4-aminobenzoate HCl) 895  
 Proclavine (3,6-Diaminoacridine) 475  
 Proflavin sulfate 426  
 Progesterone 854  
 Propacil (6-Propyl-2-thiouracil) 475  
 PROPHOS [1,2-Bis(diphenylphosphino)propane] 198, 726  
 Propylail (6-Propyl-2-thiouracil) 475  
 Proteins 812 purification 768  
 Proteoglycans 812

Proteomics 769  
 Proton ionophore (4-Nonadecylpyridine) 467  
 Proton sponge (*N,N,N',N'*-Tetramethyl-1,8-naphthalenediamine) 384  
 Protoporphyrin IX (& ester) 895  
 Protopine 475  
 Provitamin B, D-panthenol 891  
 Provocholine (*O*-Acetyl- $\beta$ -methylcholine chloride) 857  
 Proxyphylline [( $\pm$ )-7-(2-Hydroxypropyl)theophylline] 448  
 PS (Polystyrene) 377  
 PS-CIANTU 315  
 PS-DPPA 511  
 PTAD (4-Phenyl-1,2,4-triazoline-3,5-dione) 472  
 PUGNAC [*Z*-*O*-(2-Acetamido-2-deoxy-*D*-glycopyranosylidene-amino)-*N*-phenylcarbamate] 816  
 Purification, *via* derivatives 74  
   of alcohols 74  
   of aldehydes 75  
   of amines 75  
   of aromatic hydrocarbons 76  
   of carboxylic acids 77  
   of hydroperoxides 77  
   of ketones 77  
   of phenols 78  
   of phosphate esters 78  
   of phosphonate esters 78  
 Purification, general procedures for:  
   acetals 80  
   acids 80  
     carboxylic 80  
     sulfenic 81  
     sulfonic 81  
   acid chlorides 81  
   alcohols 81  
     monohydric 81  
     polyhydric 82  
   aldehydes 82  
   amides 82  
   amines 82  
   amino acids 83  
   anhydrides 83  
   carotenoids 83

esters 83  
 ethers 84  
 halides 84  
 hydrocarbons 85  
 hydrocarbons by zone melting 85  
 imides 86  
 imino compounds 86  
 ketones 86  
 macromolecules 765  
 nitriles 86  
 nitro compounds 86  
 nucleic acids 770  
 phenols 87  
 polypeptides 768  
 proteins 768, 812  
 quinones 87  
 salts (organic) 87  
   with metal ions 87  
   with organic cations 87  
   with sodium alkane sulfonates 88  
 sulfur compounds 88  
   disulfides 88  
   sulfones 88  
   sulfoxides 88  
   thioethers 88  
   thiols 88  
   thiolsulfonates (disulfoxides) 88

Purity, criteria of 79  
 question of 1

Purpald (4-Amino-3-hydrazino-5-mercapto-1,2,4-triazole) 403

Putrescine 2HCl 137

PyBOP [Benzotriazol-1-yloxy-(tripyrrolidino)phosphonium .PF<sub>6</sub>] 722

Pyromellitic acid (Benzene-1,2,4,5-tetracarboxylic) 274

Pyromellitic anhydride 274

## Q

Q-Phos (Pentaphenylferrocenyl(*tert*-butyl)phosphine) 746

Quercetin-3-rubinoside 905  
 ( $\pm$ )-Quinacrine (Atebrine) 481

Quinalizarin (1,2,5,8-Tetrahydroxy-9,10-anthraquinone) 378

Quinizarin (1,4-Dihydroxy-9,10-anthraquinone) 378

D-Quinovose (6-Deoxy-D-glucose) 820

Quisqualic acid (3-[3,5-Dioxo-1,2,4-oxadiazolin-2-yl]-L-alanine) 795

## R

Raoult's law 9

Rapid purification  
   of acetonitrile 106  
   of benzene 273  
   of *tert*-butyl alcohol 124  
   of carbon tetrachloride 130  
   of chloroform 133  
   of cyclohexane 229  
   of dichloromethane 141  
   of diethyl ether 143  
   of *N,N*-dimethyl-formamide 150  
   of dimethylsulfoxide 152  
   of 1,4-dioxane 434  
   of ethanol 156, 157  
   of ethyl acetate 157  
   of *n*-hexane 170  
   of methanol 181  
   of petroleum ether 196  
   of tetrahydrofuran 896  
   of toluene 386

Raubacine (Ajmalicine) 401

RDX (Cyclotrimethylenetrinitramine) 424, 932

Reagents and purity 67

Recrystallisation 14, 61  
   choice of solvents 16  
   from the melt 17  
   mixed solvents 17  
   techniques 14

Reichstein's G, adrenosterone 842

Reineckate salt 559

Rescinnamine (Anaprel, Apoterin) 483

Restriction enzymes 804, 813  
   endonucleases 804, 813  
   exonucleases 804

Retene (1-Methyl-7-isopropylphenanthrene) 379

Retinal 903

Retinoic acid (Vitamin A acid) 903

Retinol 840

Retinyl acetate 903

Retinyl palmitate 903

Reverse-phase thin layer chromatography 27

Reverse transcriptase 813  
 Riboflavin 904  
 Ribozymes 813  
 Rifamycin B, SV 897  
 Rimantadine (1-Adamantyl-ethylamine HCl) 222  
 Rivanol (6,9-Diamino-2-ethoxy-acridine *dl*-lactate monohydrate) 426, 814  
 RNA, ribonucleic acids 813  
 RNAi 814  
 dsRNA 814  
 mRNA 813  
 miRNA 814  
 shRNA 814  
 siRNA 814  
 Rongalite (Sodium formaldehyde sulfoxylate dihydrate) 531  
 Rose's metal 623  
 Rubenic acid (Dithioxamide) 153  
 Rubene, naphthacene (Benz[*b*]anthracene) 353  
 Ruppert's reagent (Trifluoromethyltrimethylsilane) 547  
 (+)-Rutin (Vitamin P) 905

## S

Saccharides 832  
 Safety 68  
 disclaimer 7  
 in the chemical laboratory 68  
 precautions 4, 68  
 Salcomine [*N,N'*-Bis(salicylidene)-ethylenediamine cobalt (II)] 630  
 Salol (Phenyl salicylate) 375  
 Salsalate acetate (*O*-Acetyl-salicylsalicylic acid) 261  
 Saltpetre (Potassium nitrate) 602  
 Salts, vapour pressure of saturated solutions 41, 42  
 SAM hydrochloride [(*-*)-S-Adenosyl-L-methionine chloride] 860  
 SAMs (Self-assembled monolayers) 931, 932, 938  
 Sanger's reagent (2,4-Dinitrofluorobenzene) 326  
 Sarcosine (*N*-Methylglycine) 795  
 SAXS (Small angle X-ray scattering) 932  
 Scavenger resins 95  
 Schlenk techniques 37  
 Schlippe's salt (Sodium thioantimonate) 616  
 Schrock-Hoveyda catalysts 687, 736  
 Schwartz' Reagent (Zirconocene chloride hydride) 711  
 Scopolamine, hyoscine, atropine 484, 897 (HBr)  
 SDS (Na Dodecylsulfate) 530  
 Selectophore {Bis[4-(1,1,3,3-tetramethylbutyl)phenyl]-phosphate calcium salt} 503  
 SE (Spectroscopic Ellipsometry) 909  
 SEC (Size exclusion chromatography) 926  
 SEM (Scanning Electron Microscope) 908, 920-923, 925  
 SEMCl [2-(Trimethylsilyl)-ethoxymethyl chloride] 545  
 Sephadex 25, 26, 49  
 Sepharose 26, 28, 49  
 Sephasorb HP 26  
 Septazen (4-*N*-Benzylsulfanilamide) 283  
 SES-Cl [2-(Trimethylsilyl)ethanesulfonyl chloride] 544  
*O*-Sialic acid (NANA) 857  
 SIBX (Stabilised IBX) 446  
 Sicacide 32  
 Sicapent 31  
 Silica 606  
 Silica gel 606  
 Silvan (2-Methylfuran) 459  
 Silver [bis(trifluoromethanesulfon)amide] AgNF<sub>2</sub> 730  
 Silylation, of glassware 4  
 of plasticware 4  
 solutions 4  
 SIMS (Scanning Ion Mass Spectrometry) 909  
 Sinigrin 832  
 Skatole (3-Methylindole) 459  
 Skellysolves A-L 201-202  
 SK-Naud catalysts 699  
 SLG (Single Layered Graphene) 930  
 Sodium *n*-alkylsulfates 528  
 Sodium borohydride, in purification 74

Sodium cyanate 611  
 Sodium D-gluconate 652  
 Sodium fluorosilicate 612  
 Sodium (metal) 609  
 Sodium methoxide 653  
 Sodium percarbonate 614  
 Sodium perchlorate (anhydrous) 615  
 Sodium poly( $\alpha$ -L-glutamate) 654  
 Sodium silicate solution 615  
 Solid phase synthesis (SPS) 94  
 of deoxyribo-nucleotides 95  
 of peptides (SPPS) 94, 100  
 Solvent(s) 5, 6, 68  
 for crystallisation 45  
 distribution 7, 68  
 extraction 7, 68  
 immiscible 7, 39  
 miscible 45  
 removal of acids from 7  
 removal of organic bases from 7  
 removal of phenols from 7  
 Solvents, mixed 17  
 Somatostatin 796  
 Soxhlet extraction 4, 15  
 S-Phos (2-Dicyclohexylphosphino-2',6'-dimethoxybiphenyl) 737  
 SPOS, SPPS, SPS 94, 100  
 Spectinomycin 2HCl (Actinospectacin) 897  
 Spectroscopy 33, 69, 79  
 SPM (Scattering Probe Microscope) 908  
 Squalane (2,6,10,15,19,23-Hexamethyl-tetracosane) 841  
 Squalene (all-*trans*-2,6,10,15,19,23-Hexamethyl-2,6,10,14,18,22-tetracosahexaene) 841  
 SRNF (Solvent resistant nanofilters) 933  
 STEIN (Stilbenediamine) 739  
 Stigmasterol (3 $\beta$ -Hydroxy-24-ethyl-cholesta-5,22-diene) 855  
 Stilbesterol (Diethylstilboesterol) 316  
 STM (Scanning Tunneling Microscope) 908, 909  
 Streptozocin (*N*-[Methylnitroso-carbamoyl]- $\alpha$ -D-glucosamine) 833  
 STS (Scanning Tunneling Spectroscopy) 908,  
 Sublimation 18  
 Sudan II (1,4-bis-(Butylamino)-9,10-dihydro-9,10-dioxoanthracene) 381  
 Sudan III [1-(*p*-Phenylazophenylazo)-2-naphthol] 381  
 Sudan IV [1-(4-*o*-Tolylazo-*o*-tolylazo)-2-naphthol] 381  
 Sudan orange G (2,4-Dihydroxy-

azobenzene) 317  
 Sugar of lead (Lead acetate) 640  
 Sulfacetamide (*N*-(4)-Acetylsulfanilamide) 261  
 Sulfuric acid (conc) for removal of alcohols 7  
 removal of ethers 7  
 removal of unsaturated hydrocarbons 7  
 Sulfuric acid (fuming) (oleum) 618–619  
 Supercritical CO<sub>2</sub> 96  
 Supercritical fluids (SCF) 96  
 Supelco columns 19, 23  
 SWCNTs (SWNTs) (Single Walled Carbon Nano Tubes) 920–924  
 SWCNT-(CO<sub>2</sub>H)<sub>n</sub> 925  
 SWCNT-functionalised 925–929  
 “Sympathetic Ink” (aqueous Cobaltous chloride) 572

## T

2,4,5-T (2,4,5-Trichlorophenoxyacetic acid) 931  
 T3P (Propylphosphonic anhydride) 527  
 Tartar emetic (Potassium antimonyl-tartrate) 645  
 TBA (Tri-*n*-butylamine) 210  
 TBAF (Tetra-*n*-butylammonium fluoride) 699, 720  
 TBD (1,5,7-Triazabicyclo[4.4.0]-dec-5-ene) 492  
 TBDMSCl (*tert*-Butyldimethylsilyl chloride) 505  
 TBDPSCl (*tert*-Butyldiphenylchlorosilane) 505  
 TBHDTB (Tri-*n*-butylhexadecylphosphonium bromide) 754  
 TBHP (*t*-Butyl hydroperoxide) 126  
 TBP (Tri-*n*-Butyl phosphine) 537  
 TCT (Cyanuric chloride) 424  
 TDM spray (4,4'-Tetramethyldiaminodiphenylmethane) 397  
 TED (DABCO, Triethylenediamine, (1,4-Diazabicyclo[2.2.2]-octane) 427  
 Tedicyp (*cis-cis-cis*-1,2,3,4-Tetrakis(diphenylphosphino-methyl)cyclopentane) 749  
 TEHP [(Tris(2-ethylhexyl)-phosphate) 552  
 TEM (Transmission Electron Microscope) 908, 909  
 TEMPO 2,2,6,6-Tetramethylpiperidin-yl-1-oxy 487  
 Terramycin (Oxytetracycline) 898  
 Terreic acid (2-Hydroxy-3-methyl-1,4-benzoquinone-5,6-epoxide) 382  
 TES [*N*-Tris(hydroxymethyl)methyl-2-aminomethane sulfonic acid] 72, 216  
 Testane (5 $\beta$ -Androstane) 842  
 Testosterone (17- $\beta$ -Hydroxy-androst-4-ene-3-one) 856  
 Tetracycline (& HCl) 899  
 6*R* and 6*S*-Tetrahydrobiopterin 2 HCl {6*R*- and 6*S*-2-Amino-4-hydroxy-6-[(1*R*,2*S*)-1,2-dihydroxypropyl]-5,6,7,8-tetrahydropteridine} 2HCl 899  
 Tetrahydrofolic acid 2HCl 899  
 Tetrahydropyran (Oxane) 487  
 Tetralone test 384  
 Tetraphene (1,2-Benzanthracene) 272  
 TFAH (tfacac; 1,1,1-Trifluoro-2,4-pentanedione) 214, 937  
 TGA (Thermo Gravimetric Analysis) 909, 920–922, 927, 928  
 Thallium (III) nitrate trihydrate [TTN, thallic nitrate trihydrate, Ti(NO<sub>3</sub>)<sub>3</sub>·3H<sub>2</sub>O] 620  
 Thaspic acid (1,14-Hexadecanedioic acid) 169  
 Theocin, theophylline (1,3-Dimethyl-2,6-dioxopurine) 488  
 Theophylline (1,3-Dimethyl-2,6-dioxopurine) 488  
 Theoretical discussion 65  
 Theoretical Plate 9

Thevetin A, and B (Cardenolide glycoside) 833  
 THF (Tetrahydrofuran) 486  
 THFA (5,6,7,8-Tetrahydrofolic acid) 899  
 Thiamine monophosphate 900  
 Thiazolsulfone (2-Amino-5-sulfanylthiazole) 268  
 Thimerosal (Sodium ethylmercuri-thiosalicylate) 531  
 6,8-Thioctic acid 882, 883  
 Thiosinamine (*N*-Allylthiourea) 114  
 Thiram (Tetramethylthiuram disulfide) 208  
 Thorium chloride 621  
 THPAL (Tris[*N*-alaninylmethylene]-phosphine) 934, 935  
 THPC (Tetrakis(hydroxymethyl)-phosphonium chloride) 552  
 Thycocymetin (Thiamphenicol) 900  
 Thyroxine {*O*-[3,5-Diiodo-4-oxyphenyl]-3,5-diiodo-D(-)-tyrosine} 796  
 Tiglic acid (*trans*-2,3-Dimethylacrylic acid) 209  
 Tibetolide (Pentadecanolide, 1-oxacyclohexadecan-2-one) 194  
 TIMP (Tissue inhibitor of metalloproteins) 814  
 TIRON (1,2-Dihydroxybenzene-3,5-disulfonic acid, di-Na salt) 509  
 TLG, thin layer gel filtration 20  
 TMCS (Trimethylsilyl chloride) 543  
 TMCTS (2,4,6,8-Tetramethylcyclo-tetrasiloxane) 535  
 TMED, TMEDA (*N,N,N',N'*-Tetramethylethylenediamine) 208  
 TMS (Tetramethylsilane) 535  
 TMSAN (Trimethylsilylacetone nitrile) 542  
 TMSF (Trimethylsilyl fluoride) 513  
 TMSI (Trimethylsilyliodide) 515  
 TMSNF<sub>2</sub> [Trimethylsilyl bis-(trifluoromethanesulfon)amide] 731  
 TNBS (2,4,6-Trinitrobenzenesulfonic acid) 393  
 TNT (2,4,6-Trinitrotoluene) 394, 932  
 Tolan (Diphenylacetylene) 328  
 Topitracin (Bacitracin) 862  
 TOSMIC [*p*-Tolylsulfonylmethyl isocyanide] 389  
 Toyopearl 26, 49  
 TPAP (Tetra-*n*-propylammonium perruthenate) 708, 939  
 THP (Oxane, Tetrahydropyran) 487  
 TPP 5,10,15,20-Tetraphenylporphyrin 488  
 Tranexamic acid (AMCHA) 776  
*trans*-Traumatic acid (2-Dodecene-1,12-dioic acid) 209  
 Traut's Reagent (2-Iminotetrahydrothiophene) 450  
 TREN [Tris(2-aminoethyl)amine] 216

Tricarballic acid (Propane-1,2,3-tricarboxylic acid) 210  
 TRICINE [*N*-Tris-(hydroxymethyl)methylglycine] 797  
 TRIEN (TETA) (Triethylene-tetramine, trientine) 212  
 TRIEN.HCl 213  
 Triglyme (Triethylene glycol dimethyl ether) 212  
 Trigonelline (1-Methylnicotinic acid zwitterion) 492  
 Trigonellamide chloride (1-Methyl nicotinamide chloride) 797  
 Trimellitic acid, trimesic acid (Benzene-1,3,5-tricarboxylic) 275  
 Triphosgene [Bis(trichloromethyl) carbonate] 119  
 TRIS [*N*-Tris(hydroxymethyl)methylamine] 72, 216  
 TRIS-HCl [*N*-Tris(hydroxymethyl)methyl amine HCl] 216  
 Tritium 621–622  
 Triton B (Benzyltrimethylammonium hydroxide) 283  
 Tropaeolin 0 530  
 Tropaeolin 00 553  
 Tropaeolin 000 520, 521  
 Tropine (3-Tropanol) 495  
 TTC (TTZ) (2,3,5-Triphenyltetrazolium chloride) 494  
 TTMSS [Tris(trimethylsilyl)silane] 553  
 TTN (Thallium III nitrate) 235, 620  
 Tungsten (rod) 621–622  
 Tunicamycin 901  
 Tuberactinomycin (Viomycin sulfate) 903  
 Tyrocidine A (Cyclic decapeptide) 797

## U

Ubiquinone, Coenzyme Q 868  
 UF (Ultrafiltration) 16, 767, 933

UHP (Hydrogen peroxide urea adduct) 582  
 Ulexine (Cytisine) 424  
 Ultrafiltration (UF) 16, 767, 933  
 Ultragels 26  
 Umbelliferone (7-Hydroxycoumarin) 495  
 UPLC (Ultraperformance liquid chromatography) 28  
 Uracil 496  
 Uramil (5-Aminobarbituric acid) 496  
 Urethane (Ethyl carbamate) 217  
 Uridine and uridine nucleotides 901  
 Uridine 5'-(1-thio) monophosphate 901  
 Urocanic acid (4-Imidazolylacrylic acid) 496  
 Urotropin, HMTA (Hexamethylenetetramine) 444  
 Ursocholic acid 845  
 Ursodeoxycholic acid (Ursodiol) 856  
 Ursodiol (3 $\alpha$ ,7 $\beta$ -Dihydroxy-5 $\beta$ -cholan-24-oic acid) 856  
 Usnic 902  
 Ustizeain B (Ustilagic acid, di-D-glucosyldihydroxyhexadecanoic acid) 834

## V

*cis*-Vaccenic acid (Octadec-11-enoic acid) 217  
 Vacuum lines 37  
*S*-Valinol (*S*-2-Amino-3-methyl-1-butanol) 778  
 Valinomycin 902  
 ( $\pm$ )-Verapamil HCl (5-[*N*-(3,4-Dimethoxyphenylethyl)methylamino]-2-[3,4-dimethoxyphenyl]-2-isopropylvaleronitrile HCl) 902  
 Veratridine (3-Veratroylveracevine) 902  
 Verkade Superbase (2,8,9-Trimethyl-*pro*-azaphosphatrane) 547  
 Verkade Superbase HCl 548  
 Vicine (2,4-Diamino-5- $\beta$ -D-glucopyranosidoxy-6-hydroxypyrimidine) 834

Vilsmeier Reagent [(Chloromethylene)dimethyliminium chloride] 324  
 Vinblastine sulfate (VLB) 902  
 ( $\pm$ )-Vinclozolin [3-(3,5-Dichlorophenyl)-5-methyl-5-vinylloxazolidine-2,4-dione] 496  
 Vincristine sulfate 902  
 Viocin (Viomycin sulfate) 903  
 Violanthrene (Dibenzanthrene) 396  
 Viologen (4,4'-Dipyridyl dihydrochloride) 497  
 Violuric acid (5-Isonitrosobarbituric acid) 453  
 Virofrol (1-Adamantylamine HCl) 221  
 Visnagin (4-Methoxy-7-methyl-5*H*-furo[3,2-*g*][1]benzopyran-5-one) 497  
 VPA (Valproic acid, 2-propylpentanoic Acid) 200  
 Vitamins (A, B<sub>1</sub>, B<sub>2</sub>, B<sub>3</sub>, B<sub>6</sub>, B<sub>12</sub>, D<sub>2</sub>, D<sub>3</sub>, E, K<sub>1</sub>, K<sub>3</sub>, K<sub>5</sub>, P, U) 903-905  
 VLB, vinblastine 902

## W

Wanzlick's Reagent (*N,N'*-Diphenylethylenediamine) 329  
 Warfarin (4-Hydroxy-3-(3-oxo-1-phenylbutyl)-2*H*-1-benzopyran-2-one) 906  
 Water 623  
 White vitriol (Zinc sulfate) 625  
 Wilkinson's catalyst [Chlorotris-(triphenylphosphine)rhodium(I)] 679  
 Wood's Metal 623  
 Woodward's Reagent K (*N*-Ethyl-5-phenylisoxazolinium-3'-sulfonate) 439  
 Wurster's reagent [*N,N,N',N'*-Tetramethyl-1,4-phenylenediamine dihydrochloride] 384

## X

XANES (X-ray Absorption Near Edge Structure) 934  
 9*H*-Xanthene (Dibenzopyran) 497  
 Xanthene rhamnoside 834  
 Xanthine (2,6-Dihydroxypurine) 497  
 Xanthone (9-Xanthenone) 497  
 Xanthophyll (Lutein) 841

- Xanthopterin 906  
Xanthotoxin 906  
Xanthurenic acid (5,8-Dihydroxy-quinoline-2-carboxylic acid) 498  
Xantphos [4,5-Bis(diphenylphosphino)-9,9-dimethylxanthene] 673, 710, 711, **751**  
XFS (X-ray Fluorescence Spectroscopy) 909  
X-Phos (2-Dicyclohexylphosphino-2',4',6'-triisopropylbiphenyl) 678, 679, **738**  
XPS (X-ray Photoelectron Spectroscopy) 909, 927, 936, 937  
Xylenol Orange 498  
Xylyl-P-Phos (2,2',6,6'-Tetra-methoxy-4,4'-bis[di(3,5-xylyl)phosphino]-3,3'-bipyridine) 677, 707, **749**

## Y

- $\alpha$ -Yohimbine 401, **498**  
 $\gamma$ -Yohimbine 401, 498

## Z

- Zeatin (*trans*-*N*<sup>6</sup>-[4-Hydroxy-3-methylbut-2-en-1-yl]adenine) 906  
Zeaxanthin [all*trans*- $\beta$ -Carotene-3,3'(*R,R'*)-diol] 841  
Zeinoxanthin (all-*trans*-3*R*,6*R'*- $\beta$ ,*e*-Carotene-3-ol) 839  
Zinc-copper alloy 934  
copper/iron oxide 934  
couple 623  
Zinc-Copper Couple 624  
Zinc nano metallic powder 938  
Zinc-Silver couple 624  
Zincon {*o*-[1-(2-Hydroxy-5-sulfo)-3-phenyl-5-formazono]-benzoic acid} 661

## CAS Registry Numbers Index

### HOW TO VALIDATE CHEMICAL ABSTRACTS SERVICE REGISTRY NUMBERS

**Almost** all chemical and biochemical entries in this book have CAS (Chemical Abstract Service) Registry Numbers to identify them and have been entered for each substance. The registry numbers are an efficient way for identifying substances and are independent of the nomenclature of substances. Unlike chemical names which may have more than one synonymous name, there is only one CAS Registry Number for each substance (with only a few exceptions, e.g. where a substance may have another number before purification, or before determination of absolute configuration).

The registry numbers are made up of three sets of digits separated by two hyphens with the general form:  $nN_n + \dots + 4N_4 + 3N_3 + 2N_2 + 1N_1 - R$  where the  $N$ s are sequential numbers and  $R$  is the *check digit*. The first set can be up to eight digits (at this point in time, but not less than two), and the second is made of two digits (zeros are included as a digits) and the third is a single digit which *checks* the validity of the sequential numbers, e.g. 379669-72-4 for *but-3-enylboronic acid* or 83-32-9 for *acenaphthene*. The validity of a registry number can be checked by the following formula (computer generated):

$$\frac{nN_n + \dots + 4N_4 + 3N_3 + 2N_2 + 1N_1}{10} = Q + R/10$$

where the  $N$ s include all numbers except the last *check digit*,  $Q$  represents an integer (fractions not included) which, although it is disregarded, does give the value of  $R$  equal to the *check digit* in a **valid** Registry Number.

The following three examples demonstrate the application of the formula and the validity of the Registry Number:

1. **1,5-Dihydroxynaphthalene 83-56-7**

$$\frac{(4 \times 8) + (3 \times 3) + (2 \times 5) + (1 \times 6)}{10} = \frac{57}{10} = 5 + \frac{7}{10} \quad [\text{Note: the last (check) digit of the CAS number 7 is not included.}]$$

When the value of  $Q$  is made to equal 5, the *check digit*  $R$  becomes equal to 7; thus the Registry Number is **valid**. In this case the value of  $Q$  is in fact the digit(s) (**5**), before the last *check digit* (**7**).

2. **4,5-Diamino-6-hydroxypyrimidine hemisulfate 102783-18-6**

$$\frac{(8 \times 1) + (7 \times 0) + (6 \times 2) + (5 \times 7) + (4 \times 8) + (3 \times 3) + (2 \times 1) + (1 \times 8)}{10} = \frac{106}{10} = 10 + \frac{6}{10}$$

When the value of  $Q$  is made to 10, the *check digit*  $R$  becomes equal to 6; thus the Registry Number is **valid**. In this case the value of  $Q$  is in fact the digit(s) (**10**), before the last *check digit* (**6**).

3. 1,4-Bis(pentafluorobenzyl)[C<sub>60</sub>]fullerene 1260376-31-5

$$\frac{(9 \times 1) + (8 \times 2) + (7 \times 6) + (6 \times 0) + (5 \times 3) + (4 \times 7) + (3 \times 6) + (2 \times 3) + (1 \times 1)}{10} = \frac{135}{10} = 13 + \frac{5}{10}$$

When the value of Q is made to 13, the check digit R becomes equal to 5; thus the Registry Number is **valid**. In this case the value of Q is in fact the digit(s) (**13**), before the last *check* digit (**5**).

The CAS Registry Handbook (Number Section) and supplements should be consulted for obtaining the substance name and elemental formula corresponding to a Registry Number.

To facilitate the method for locating the purification of a substance, a CAS Registry Number Index of the numbers of the entries with their respective page numbers is included here. This will also provide the reader with a rapid way to see if the purification of a particular substance has been reported in the book.

45-19-0	498	52-89-1	783	57-09-0	131, 753
50-00-0	133, 164	52-90-4	783	57-10-3	169
50-01-1	167	53-03-2	853	57-11-4	202
50-02-2	848	53-06-5	848	57-13-6	217
50-03-3	851	53-16-7	850	57-14-7	151
50-04-4	848	53-19-0	302	57-24-9	485
50-07-7	885	53-41-8	842	57-27-2	464
50-14-6	904	53-59-8	888	57-41-0	435
50-22-6	848	53-70-3	309	57-44-3	430
50-23-7	848, 851	53-84-9	887	57-47-6	438
50-27-1	850	53-96-3	258	57-48-7	822
50-28-2	850	54-06-8	400	57-50-1	833
50-29-3	285	54-11-5	888	57-53-4	179
50-30-6	313	54-28-4	901	57-55-6	197
50-32-8	277	54-47-7	896	57-57-8	468
50-33-9	471	54-62-6	861	57-62-5	862
50-36-2	867	54-64-8	531, 651	57-63-6	850
50-45-3	313	54-85-3	453	57-67-0	381
50-50-0	850	54-92-2	453	57-68-1	485
50-55-5	483	54-95-5	464	57-83-0	854
50-56-6	793	54-96-6	427	57-85-2	856
50-69-1	832	55-21-0	272	57-87-4	850
50-70-4	833	55-22-1	453	57-88-5	847
50-71-5	402	55-55-0	348	57-91-0	850
50-76-0	859	56-03-1	118	57-94-3	495
50-78-2	261	56-04-2	463	57-97-6	322
50-79-3	140, 313	56-06-4	426	58-00-4	862
50-81-7	819	56-10-0	115	58-08-2	418
50-84-0	313	56-12-2	776	58-15-1	406
50-85-1	305	56-17-7	782	58-18-4	853
50-89-5	492	56-18-8	137	58-22-0	856
50-98-6	875	56-23-5	130	58-27-5	905
50-99-7	825	56-23-1	316	58-46-8	486
51-05-8	895	56-24-6	659	58-55-9	488
51-17-2	409	56-34-8	205	58-56-0	904
51-21-8	440	56-40-6	774	58-58-2	896
51-28-5	326	56-41-7	774	58-60-6	897
51-34-3	484	56-45-1	796	58-61-7	400
51-35-4	790	56-47-3	848	58-63-9	882
51-36-5	313	56-54-2	482, 712	58-64-0	859
51-41-2	889	56-55-3	272	58-85-5	863
51-43-4	860	56-57-5	467	58-86-6	834
51-44-5	313	56-65-5	860	58-89-9	249
51-45-6	444	56-75-7	866	58-90-2	382
51-48-9	796	56-81-5	167	58-95-7	906
51-49-0	796	56-82-6	167	58-96-8	496
51-52-5	475	56-84-8	780	59-00-7	498
51-55-8	395, 407	56-85-9	786	59-02-9	905
51-60-5	357	56-86-0	786	59-05-2	860
51-67-2	395	56-87-1	791	59-14-3	864
51-74-1	444	56-88-2	783	59-23-4	823
51-78-5	267	56-89-3	784	59-30-3	877
51-79-6	217	56-91-7	776, 778	59-31-4	449
52-21-1	853	56-92-8	444	59-46-1	377
52-39-1	842	56-93-9	283	59-49-4	410
52-52-8	776	57-00-1	782	59-50-7	299
52-67-5	891	57-06-7	114	59-51-8	792

59-52-9	147	63-89-8	874	68-81-9	233
59-53-0	858	63-91-2	712, 794	68-81-5	452
59-56-3	879	63-92-3	369	68-94-0	448
59-66-5	397	64-04-0	367	69-05-6	481
59-67-6	466	64-10-8	376	69-33-0	901
59-88-1	373	64-17-5	2, 30, 156	69-57-8	862
59-92-7	784	64-18-6	165	69-65-8	828
59-96-1	369	64-19-7	105	69-72-7	380
59-99-4	357	64-20-0	207	69-81-8	401
60-00-4	161	64-67-5	146	69-89-6	497
60-09-3	262	64-69-7	173	69-93-2	496
60-10-6	331	64-72-2	862	70-11-1	287
60-11-7	320	64-73-3	871	70-18-8	786
60-12-8	372	64-75-5	899	70-23-5	158
60-18-4	797	64-85-7	848	70-26-8	793
60-19-5	395	64-86-8	819, 868	70-30-4	337
60-20-8	788	65-22-5	896	70-34-8	326
60-23-1	783	65-45-2	379	70-47-3	779
60-24-2	179	65-46-3	870	70-49-5	209
60-27-5	782, 880	65-49-6	267	70-55-3	387
60-29-7	143	65-71-4	492	70-69-9	267
60-31-1	857	65-82-7	773	70-70-2	342
60-34-4	185	65-85-0	276	70-78-0	791
60-35-5	104	65-86-1	467	71-00-1	788
60-46-8	785	66-22-8	496	71-23-8	199
60-54-8	899	66-23-9	857	71-30-7	871
60-56-0	457	66-27-3	185	71-36-3	122
60-80-0	407	66-71-7	470	71-41-0	116
60-81-1	831	66-77-3	353	71-43-2	272
60-82-2	376	66-81-9	424	71-55-6	211
61-25-6	468	66-84-2	825	71-58-9	852
61-33-6	863	66-99-9	353	71-62-5	902
61-54-1	495	67-03-8	903	71-63-6	821
61-72-3	346	67-42-5	161	71-67-0	504
61-73-4	458	67-43-6	143	71-91-0	205
61-78-2	284	67-47-0	447	72-12-0	204
61-78-9	777	67-48-1	134	72-18-4	797
61-82-5	406	67-51-6	433	72-19-5	796
61-90-5	791	67-52-7	408	72-40-2	404
62-13-5	860	67-56-1	180	72-43-5	347
62-23-7	360	67-63-0	175	72-48-0	401
62-38-4	644	67-64-1	105	72-54-8	285
62-44-2	259	67-66-3	133	72-55-9	285
62-51-1	857	67-68-5	152	72-80-0	429
62-53-3	268	67-71-0	152	73-03-0	871
62-54-4	631	67-72-1	168	73-22-3	797
62-55-5	209	67-96-9	849	73-24-5	400
62-56-6	209	67-97-0	905	73-32-5	791
62-57-7	777	67-99-2	879	74-11-3	298
62-74-8	651	68-04-2	660	74-31-7	330
62-75-9	151	68-05-3	205	74-39-5	359
62-76-0	653	68-11-1	209	74-79-3	779
63-05-8	842	68-12-2	150	74-82-8	179
63-42-3	827	68-19-9	904	74-83-9	182
63-45-6	475	68-26-8	903	74-84-0	155
63-68-3	792	68-41-7	782	74-85-1	160

74-86-2	109	75-86-5	106	78-59-1	251
74-87-3	183	75-87-6	131	78-61-1	117, 915
74-88-4	185	75-89-8	214	78-71-7	412
74-89-5	182	75-91-2	126	78-76-2	119
74-90-8	582	75-94-5	553	78-77-3	174
74-96-4	158	75-97-8	128	78-78-4	182
74-98-6	197	75-98-9	197	78-79-5	175
74-99-7	201	76-01-7	194	78-82-0	175
75-00-3	159	76-03-9	210	78-83-1	187
75-03-6	162	76-05-1	213	78-84-2	174
75-04-7	158	76-06-2	133	78-85-3	179
75-05-8	106	76-09-5	196, 747	78-86-4	132
75-07-0	104	76-13-1	211	78-87-5	141
75-08-1	156	76-16-4	169	78-90-0	197
75-09-2	141	76-19-7	192	78-92-2	124
75-11-6	147	76-24-4	402	78-93-0	123
75-12-7	165	76-32-4	225	78-94-4	189
75-15-0	568	76-37-9	206	78-95-5	132
75-18-3	152, 732	76-39-1	186	78-98-8	184
75-21-8	439	76-57-3	867	79-00-5	211
75-22-9	504	76-59-5	292	79-01-6	211
75-24-1	627	76-60-8	289	79-05-0	198
75-25-2	119	76-73-3	484	79-06-1	111
75-26-3	176	76-74-4	469	79-07-2	132
75-28-5	174	76-83-5	394	79-08-3	119
75-29-6	176	76-84-6	394	79-09-4	199, 200
75-30-9	177	76-86-8	507	79-10-7	111
75-33-2	198	76-87-9	659	79-11-8	132
75-34-3	140	76-93-7	275	79-14-1	167
75-36-5	109	77-06-5	249	79-15-2	119
75-44-5	596	77-09-8	368	79-16-3	181
75-47-8	173	77-21-4	439	79-19-6	209
75-50-3	214	77-26-9	402	79-20-9	182
75-52-5	190	77-47-4	249	79-24-3	190
75-54-7	517	77-48-5	428	79-27-6	203
75-56-9	475	77-49-6	186	79-29-8	149
75-57-0	207	77-53-2	226	79-31-2	175
75-58-1	207	77-59-8	901	79-33-4	177
75-59-2	207	77-71-4	433	79-34-5	204
75-60-5	505	77-73-6	235, 634	79-38-9	134
75-62-7	121	77-85-0	216	79-39-0	179
75-63-8	121	77-86-1	216	79-40-3	153
75-64-9	124	77-92-9	135	79-41-4	179
75-65-0	124	77-93-0	135	79-43-6	140
75-66-1	187	77-95-2	255	79-44-7	149
75-71-8	140	77-99-6	215	79-46-9	190
75-72-9	134	78-00-2	655	79-49-2	194
75-73-0	130	78-09-1	206	79-57-2	898
75-75-2	180	78-10-4	533	79-58-3	854
75-75-2	237	78-11-5	831	79-63-0	852
75-76-3	535, 545	78-27-3	248	79-69-6	250
75-77-4	543, 687, 694	78-28-4	187	79-74-3	308
75-78-5	510	78-40-0	539	79-77-6	250
75-79-6	518	78-42-2	552	79-81-2	903
75-83-2	149	78-46-6	508	80-05-7	345
75-84-3	152	78-50-2	548	80-11-5	138, 352
75-85-4	116	78-53-5	509	80-15-9	305

80-43-3	315	83-67-0	488	86-88-4	357
80-46-6	268	83-72-7	341	86-93-1	471
80-48-8	353	83-73-8	431	86-98-6	430
80-59-1	209	83-74-9	449	87-13-8	144
80-62-6	185, 233	83-88-5	904	87-17-2	379
80-70-6	208	84-11-7	367	87-32-1	774
80-72-8	255	84-12-8	470	87-41-2	376
80-73-9	433, 916	84-15-1	381	87-42-3	421
80-78-4	854	84-16-2	285	87-48-9	415
80-91-1	854	84-21-9	859	87-51-4	451
80-92-2	854	84-24-2	376	87-52-5	443
80-97-7	846	84-54-8	349	87-56-9	189
81-07-2	484	84-58-2	314	87-59-2	321
81-08-3	485	84-65-1	270	87-61-6	391
81-13-0	891	84-66-2	316	87-62-7	322, 324
81-14-1	345	84-68-4	312	87-68-3	168
81-16-3	356	84-74-2	311	87-72-9	818
81-24-3	855	84-76-4	328	87-81-0	833
81-25-4	847	84-79-7	345	87-85-4	249, 337
81-27-6	832	84-80-0	905	87-86-5	366
81-31-2	396	84-86-6	355	87-87-6	317
81-54-9	378	84-88-8	449	87-88-7	296
81-61-8	378	84-89-9	356	87-89-8	882
81-64-1	378, 381	85-01-8	367	88-04-0	298
81-77-6	450	85-02-9	409	88-06-2	391
81-81-2	906	85-18-7	422	88-14-2	441
81-84-5	354	85-31-4	900	88-15-3	399
81-88-9	483	85-32-5	443	88-16-4	298, 419
81-92-5	369	85-41-6	376	88-19-7	386
81-93-6	470	85-42-7	230	88-21-1	365
82-02-0	454	85-44-9	376	88-27-7	311
82-05-3	272	85-46-1	354	88-30-2	391
82-22-4	270	85-47-2	354	88-31-1	491
82-24-6	262	85-52-9	278	88-58-4	311
82-38-2	348	85-61-0	867	88-65-3	287
82-57-5	497	85-66-5	476	88-67-5	343, 445
82-58-6	456	85-83-6	381	88-72-2	364
82-62-2	391	85-84-7	273	88-73-3	300
82-68-8	366	85-85-8	479	88-74-4	358
82-71-3	380	85-86-9	381	88-75-5	362
82-86-0	258	86-00-0	263, 360	88-88-0	377
82-94-0	350	86-28-2	438	88-89-1	377
83-07-8	402	86-29-3	328	88-99-3	376
83-12-5	373	86-34-0	369	89-01-0	476
83-30-7	392	86-35-1	439	89-05-4	274
83-32-9	258	86-38-4	429	89-25-8	460
83-34-1	459	86-42-0	407	89-32-7	274
83-38-5	312	86-48-6	355	89-51-0	338
83-40-9	305	86-54-4	445	89-52-1	259
83-43-2	852	86-55-5	355	89-56-5	305
83-44-3	848	86-57-7	361	89-57-6	267
83-46-5	854	86-59-9	482	89-61-2	314
83-48-7	855	86-73-7	334	89-63-4	300
83-49-8	852	86-74-8	418	89-65-6	827
83-56-7	318	86-84-0	357	89-69-0	391
83-60-3	483	86-87-3	356	89-73-6	380

89-82-7	255	91-58-7	300	93-62-9	172
89-83-8	386, 760	91-59-8	356	93-76-5	391
89-86-1	317	91-60-1	385	93-91-4	277
89-91-2	146	91-62-3	462	93-95-8	373
89-98-5	297	91-63-4	462	93-97-0	276
90-00-6	334	91-64-5	423	93-98-1	272
90-01-7	339	91-65-6	246	93-99-2	371
90-02-8	379	91-66-7	316	94-09-7	333
90-04-0	269	91-94-1	312	94-20-2	377
90-05-1	337	92-06-8	382	94-25-7	293
90-11-9	291	92-24-0	353	94-36-0	278
90-12-0	350, 913	92-29-5	338	94-41-7	271
90-13-1	300	92-32-0	481	94-53-1	377
90-15-3	355	92-35-3	489	94-59-7	484
90-19-7	483	92-41-1	354	94-62-2	474
90-20-0	266	92-51-3	223	94-67-7	379
90-26-6	371	92-52-4	284	94-74-6	299
90-30-2	374	92-59-1	281	94-75-7	315
90-33-5	446	92-61-5	484	95-03-4	299
90-41-5	263	92-62-6	475	95-13-6	265, 343
90-43-7	339	92-66-0	289	95-14-7	410
90-44-8	270	92-67-1	263	95-15-8	410
90-45-9	402	92-69-3	339	95-20-5	459
90-46-0	498	92-70-6	355	95-45-4	151
90-47-1	497	92-71-7	435	95-46-5	292
90-52-8	404	92-76-2	341	95-47-6	396
90-65-3	892	92-77-3	341	95-48-7	304
90-80-2	824	92-82-0	470	95-49-8	302
90-82-4	377	92-83-1	497	95-50-1	312
90-90-4	288	92-84-2	471	95-51-2	296
90-91-5	900	92-85-3	489	95-52-3	336
90-93-7	285	92-86-4	310	95-53-4	388
90-94-8	353	92-87-5	275	95-54-5	372
90-96-0	319	92-88-6	318	95-55-6	267
90-98-2	313	92-91-1	370	95-56-7	291
90-99-3	330	92-92-2	284	95-57-8	301
91-00-9	330	92-94-4	382	95-63-6	392
91-01-0	275	93-00-5	356	95-64-7	322
91-02-1	410	93-04-9	351	95-65-8	293, 323
91-08-7	389	93-05-0	263	95-68-1	321
91-10-1	319	93-07-2	319	95-71-6	387
91-13-4	310	93-08-3	260	95-73-8	315
91-15-6	376	93-09-4	355	95-75-0	315
91-16-7	318	93-10-7	481	95-76-1	312
91-17-8	241	93-11-8	354	95-77-2	314
91-18-9	475	93-15-2	334	95-78-3	321
91-19-0	483	93-17-4	338	95-80-7	386
91-20-3	353	93-18-5	333	95-83-0	302
91-22-5	482	93-20-9	357	95-84-1	264
91-23-6	358	93-25-4	347	95-87-4	323
91-40-7	328	93-31-9	492	95-88-5	302
91-47-4	372	93-35-6	495	95-93-2	384
91-48-5	372	93-40-3	320	95-94-3	382
91-52-1	319	93-53-8	374	95-95-4	391
91-53-2	438	93-56-1	380	95-96-5	432
91-56-5	452	93-58-3	349	96-09-3	380
91-57-6	351	93-59-4	367	96-10-6	636

96-14-0	186	98-58-8	287	100-09-4	347
96-20-8	115	98-59-9	387	100-10-7	320
96-22-0	145	98-60-2	297	100-11-8	360
96-31-1	153	98-74-8	359	100-13-0	364
96-33-3	182	98-79-3	795	100-14-1	360
96-34-4	183	98-80-6	521, 526	100-15-2	351
96-37-7	253	98-82-8	305	100-16-3	362
96-41-3	240, 241	98-83-9	352	100-17-4	358
96-43-5	422	98-85-1	372	100-19-6	358
96-45-7	439	98-86-2	259	100-20-9	381
96-47-9	463	98-88-4	278	100-21-0	381
96-48-0	129	98-89-5	230	100-22-1	384
96-49-1	434	98-92-0	465	100-23-2	323
96-50-4	406	98-95-3	358	100-25-4	325
96-53-7	489	98-96-4	476	100-26-5	478
96-54-8	462	98-97-5	476	100-28-7	363
96-77-5	368	98-98-6	473	100-39-0	279
96-82-2	827	99-03-6	262	100-42-5	380
96-97-9	363	99-04-7	388	100-43-6	497
96-98-0	352	99-05-8	262	100-44-7	280
97-00-7	298	99-06-9	339	100-46-9	279
97-02-9	325	99-08-1	364	100-47-0	276
97-23-4	314	99-09-2	358	100-48-1	424
97-30-3	829	99-11-6	423	100-51-6	279
97-39-2	332	99-14-9	210	100-52-7	271
97-41-6	247	99-24-1	350	100-53-8	281
97-44-9	499	99-28-5	310	100-54-9	424
97-53-0	334	99-30-9	314	100-57-2	644
97-56-3	262	99-32-1	419	100-61-8	348
97-59-6	401	99-33-2	326	100-63-0	373
97-62-1	162	99-34-3	326	100-64-1	232
97-63-2	162	99-35-4	393	100-65-2	373
97-65-4	177	99-52-5	351	100-66-3	269
97-67-6	178	99-54-7	314	100-69-6	497
97-88-1	127	99-55-8	351	100-71-0	439
97-90-5	161	99-56-9	362	100-76-5	483
97-93-8	626	99-57-0	266	100-83-4	339
97-94-9	539	99-61-6	358	100-85-6	283
97-95-0	159	99-65-0	325	100-97-0	444
97-97-2	132	99-66-1	200	101-01-9	394
98-00-0	442	99-73-0	291	101-02-0	550
98-01-1	442	99-76-3	350	101-31-5	395
98-03-3	490	99-77-4	334	101-43-9	233
98-05-5	521	99-81-0	291	101-49-5	411
98-06-6	293	99-87-6	306	101-53-1	282
98-08-8	392	99-90-1	287	101-54-2	264
98-09-9	274	99-92-3	262	101-61-1	384
98-11-3	274	99-93-4	339	101-63-3	286
98-15-7	298	99-94-5	388	101-64-4	348
98-29-3	293	99-96-7	339	101-75-7	273
98-48-6	273	99-97-8	325	101-77-9	308
98-50-0	500	99-99-0	364	101-80-4	307, 937
98-51-1	294	100-00-5	301	101-81-5	330
98-53-3	224	100-01-6	358	101-82-6	411
98-54-4	294	100-02-7	362	101-83-7	246
98-56-6	298	100-06-1	346	101-84-8	373

101-97-3	334	104-91-6	364	107-05-1	113
102-01-2	259	104-92-7	287	107-06-2	140
102-02-3	371	104-93-8	349	107-07-3	132
102-04-5	309	104-94-9	269	107-08-4	200
102-06-7	330	104-96-1	350	107-10-8	199
102-08-9	331	104-98-3	496	107-11-9	113
102-09-0	329	105-30-6	187	107-12-0	199
102-25-0	391	105-34-0	183	107-13-1	111
102-54-5	638	105-36-2	158	107-14-2	132
102-69-2	215	105-37-3	163	107-15-3	144, 150, 160
102-71-6	211	105-39-5	159	107-18-6	113
102-82-9	210	105-41-9	115	107-19-7	198
102-85-2	538	105-53-3	145	107-21-1	161
102-87-4	211, 755	105-54-4	159	107-27-7	637
102-92-1	303	105-55-5	146	107-30-2	133
102-94-3	303	105-56-6	160	107-31-3	184
103-03-7	375	105-57-7	104	107-40-4	215
103-16-2	282	105-58-8	143	107-41-5	186, 187
103-19-5	389	105-59-9	184	107-43-7	780
103-29-7	283, 544	105-60-2	418	107-46-0	514
103-30-0	283, 380	105-61-3	178	107-51-7	520
103-32-2	279	105-74-8	178	107-59-5	125
103-33-3	271	105-87-3	878	107-71-1	128
103-34-4	437	106-02-5	194	107-81-3	120
103-36-6	333	106-22-9	135	107-82-4	173
103-41-3	280	106-24-1	166, 878	107-83-5	186
103-49-1	309	106-28-5	876	107-84-6	173
103-50-4	281	106-31-0	129	107-86-8	201
103-63-9	367, 471	106-34-3	378	107-87-9	188
103-70-8	336	106-36-5	201	107-89-1	112
103-71-9	373	106-37-6	310	107-91-5	135
103-72-0	374	106-38-7	292	107-92-6	129
103-73-1	368	106-40-1	287	107-93-7	135
103-79-9	370	106-41-2	291	107-94-8	134
103-81-1	369	106-42-3	396	107-95-9	774
103-82-2	369	106-43-4	302	107-97-1	795
103-83-3	280	106-44-5	304	107-98-2	200
103-84-4	259	106-45-6	387	108-01-0	148
103-85-5	375	106-46-7	312	108-03-2	190
103-86-6	267	106-47-8	296	108-05-4	218
103-88-8	286	106-48-9	301	108-08-7	152
103-89-9	259	106-49-0	388	108-10-1	187
103-90-2	258	106-50-3	372	108-11-2	187
104-01-8	347	106-51-4	277	108-13-4	178
104-03-0	362	106-54-7	302	108-18-9	147
104-04-1	358	106-57-0	786	108-19-0	119
104-12-1	302	106-70-7	185	108-20-3	176
104-24-5	370	106-88-7	155	108-21-4	176
104-51-8	293	106-89-8	155, 332	108-22-5	175, 432
104-54-1	304	106-93-4	138	108-24-7	105
104-74-5	437	106-94-5	199	108-29-2	496
104-81-4	292	106-95-6	113	108-30-5	202
104-82-5	349	106-97-8	121	108-31-6	442
104-83-6	298	106-99-0	121	108-32-7	475
104-85-8	389	107-02-8	110	108-38-3	396
104-87-0	386	107-03-9	198	108-39-4	304
104-88-1	297	107-04-0	132	108-41-8	302

108-43-0	301	109-79-5	122	110-99-6	146
108-44-1	388	109-82-0	184	111-01-3	841
108-45-2	372	109-85-3	181	111-02-4	841
108-46-3	379	109-86-4	181	111-11-5	186
108-47-4	455	109-87-5	165, 695	111-15-9	108
108-48-5	456	109-89-7	142	111-16-0	196
108-52-1	405	109-90-0	162	111-17-1	209
108-64-5	162	109-92-2	164	111-20-6	201
108-65-6	200	109-94-4	162	111-25-1	171
108-67-8	346	109-97-7	481	111-26-2	171
108-68-9	324	109-99-9	486	111-27-3	170
108-69-0	322	110-00-9	441	111-30-8	166
108-70-3	391	110-01-0	487	111-34-2	128
108-73-6	376	110-02-1	490	111-40-0	137
108-75-8	493, 744	110-03-2	151	111-41-1	172
108-77-0	424	110-05-4	140	111-42-2	142
108-78-1	457	110-13-4	107	111-43-3	200
108-79-2	433	110-14-5	202	111-44-4	118
108-80-5	424	110-15-6	202	111-46-6	143
108-85-0	233	110-16-7	178	111-47-7	153
108-86-1	287	110-17-8	166	111-48-8	678, 679, 726
108-87-2	252	110-18-9	208	111-49-9	444
108-88-3	386	110-20-3	106	111-55-7	161
108-89-4	461	110-21-4	171	111-61-5	163
108-90-7	297	110-26-9	118	111-63-7	218
108-91-8	233	110-42-9	184	111-65-9	192
108-93-0	231	110-44-1	202	111-66-0	192
108-94-1	232	110-45-2	175	111-68-2	168
108-95-2	368	110-49-6	108	111-69-3	112
108-97-4	480	110-51-0	504	111-70-6	168
108-98-5	385	110-53-2	116	111-71-7	167
108-99-6	461	110-54-3	170	111-76-2	123
109-00-2	448	110-58-7	116	111-77-3	143
109-02-4	459	110-59-8	218	111-78-4	234, 629
109-04-6	416	110-61-2	203	111-78-4	634, 684
109-06-8	461	110-62-3	218	111-82-0	184
109-07-9	460	110-63-4	121	111-83-1	193
109-08-0	461	110-65-6	129	111-84-2	190
109-09-1	421	110-66-7	195	111-86-4	192
109-12-6	406	110-70-3	144, 149, 718	111-87-5	192
109-13-7	128	110-71-4	147	111-88-6	192
109-52-4	218	110-74-7	200	111-90-0	143
109-53-5	174	110-75-8	132	111-92-2	139
109-57-9	114	110-80-5	157	111-96-6	146, 243
109-60-4	199	110-81-6	143	112-00-5	155, 754
109-63-7	717	110-82-7	229	112-02-7	131, 753
109-64-8	139	110-83-8	232, 235	112-04-9	520
109-65-9	125	110-85-0	473	112-05-0	192, 194
109-66-0	195	110-86-1	477	112-07-2	108
109-68-2	195	110-87-2	430	112-24-3	212, 216
109-69-3	125	110-88-3	494	112-27-6	212
109-70-6	134	110-89-4	474	112-29-8	136
109-73-9	124	110-91-8	464	112-30-1	136
109-74-0	129	110-94-1	166	112-34-5	143
109-75-1	113	110-97-4	147	112-36-7	143
109-77-3	178	110-98-5	153	112-38-9	217

112-40-3	154	117-08-8	383	120-80-9	296
112-42-5	217	117-10-2	317	120-82-1	391
112-47-0	136	117-12-4	270	120-83-2	314
112-48-1	162	117-18-0	382	120-92-3	240
112-49-2	212	117-34-0	328	120-93-4	439
112-53-8	154	117-37-3	269	121-00-6	294
112-54-9	178	117-39-3	481	121-14-2	327
112-55-0	154	117-80-6	314	121-19-7	520
112-57-2	206	117-81-7	316	121-33-5	395
112-62-9	186	118-00-3	443	121-40-4	382
112-72-1	205	118-10-5	712, 753, 867	121-43-7	541
112-73-2	139	118-48-9	452	121-44-8	212
112-78-8	217	118-52-5	429	121-45-9	542
112-79-8	155	118-55-8	375	121-46-0	253, 745
112-80-1	193	118-65-0	250	121-47-1	346
112-82-3	131, 754	118-69-4	315	121-51-7	359
112-85-6	154	118-71-8	456	121-52-8	359
112-86-7	155	118-74-1	337	121-54-0	275
112-89-0	120	118-75-2	296	121-57-3	381
112-92-5	191	118-78-5	496	121-60-8	258
112-95-8	155	118-79-6	390	121-69-7	321
113-00-8	167	118-90-1	387	121-70-0	522
113-15-5	438	118-91-2	297	121-72-2	324
113-58-4	863	118-92-3	262, 343	121-73-3	301
113-73-5	788	118-96-7	394	121-79-9	377
114-07-8	876	118-97-8	298	121-82-4	424
114-21-6	652	119-24-4	476	121-88-0	266
114-25-0	412	119-26-6	327	121-89-1	358
114-49-8	897	119-27-7	325	121-91-5	344, 446
114-70-5	653	119-30-2	344	121-92-6	360
114-76-1	653	119-32-4	351	121-98-2	348
114-80-7	357	119-36-8	352	122-00-9	348
114-83-0	261	119-39-1	472	122-03-2	305
115-02-6	780	119-48-8	906	122-04-3	360
115-07-1	198	119-52-8	220	122-19-0	280
115-08-2	209	119-53-9	276	122-25-8	350
115-09-3	643	119-59-5	308	122-37-2	340
115-10-6	184	119-61-9	276	122-39-4	328
115-11-7	174	119-64-2	383	122-51-0	134, 213
115-25-3	254	119-65-3	453	122-52-1	540
115-27-5	226	119-68-6	348	122-57-6	271
115-37-7	488	119-91-5	412	122-59-8	369
115-39-9	291	119-93-7	386	122-61-8	188
115-40-2	289	120-12-7	269	122-62-3	118
115-41-3	527	120-14-9	395	122-66-7	330
115-53-7	484	120-18-3	354	122-79-2	369
115-69-5	115	120-20-7	338	122-80-5	261
115-76-4	145	120-23-0	355	122-85-0	258
115-77-5	831	120-29-6	495	122-88-3	301
115-85-5	507	120-36-5	315	123-07-9	334
115-86-6	549	120-46-7	309	123-08-0	339
115-87-7	502	120-57-0	377	123-11-5	269
116-15-4	145, 169	120-61-6	324	123-19-3	153
116-17-6	541	120-66-1	259	123-25-1	146
116-31-4	903	120-72-9	450	123-30-8	267
116-63-2	266	120-73-0	476	123-31-9	339
116-85-8	264	120-78-8	437	123-32-0	433

123-33-1	178	126-71-6	540	132-60-5	472
123-38-6	198	126-73-8	537	132-64-9	428
123-39-7	184	126-81-8	247	132-65-0	428
123-42-2	172	126-98-7	179	132-66-1	357
123-46-6	797	127-06-0	107	133-11-9	370
123-51-3	173	127-07-1	583	133-32-4	451
123-54-6	108, 639	127-09-3	649	133-38-0	146
123-56-8	202	127-17-3	201	133-49-3	366
123-62-6	199	127-18-4	204	134-32-7	356
123-63-7	468	127-19-5	148, 937	134-58-7	407
123-68-3	341	127-40-2	841	134-81-6	275, 739
123-73-9	135	127-41-3	250	134-84-9	349
123-75-1	481	127-47-9	903	134-85-0	298
123-78-4	897	127-63-9	331	134-96-3	381
123-86-4	123	127-82-2	554	135-16-0	899
123-90-0	490	127-95-7	645	135-19-3	355, 712
123-91-1	434	127-96-8	646	135-20-6	507
123-92-2	173	128-00-7	648	135-48-8	366
123-93-3	209	128-04-1	530	135-49-9	400
123-95-5	128	128-08-5	121	135-52-4	661
123-99-9	117	128-13-2	856	135-59-1	490
124-02-7	137	128-37-0	311	135-61-5	355
124-04-9	112	128-39-2	311	135-67-1	471
124-09-4	170	128-57-4	832	135-68-2	263
124-12-9	129	128-64-3	345	135-70-6	378
124-18-5	136	128-95-0	307	135-73-9	374
124-20-9	202	129-00-0	378	135-88-6	374
124-22-1	154	129-16-8	643	136-09-4	867
124-30-1	191	129-44-2	307	136-36-6	308
124-38-9	567	129-64-6	253	136-64-8	341
124-40-3	148	129-66-8	394	136-72-1	377
124-41-4	653	129-73-7	346	136-95-8	402
124-42-5	104	129-79-3	394	137-08-6	891
124-43-6	582	130-14-3	531	137-18-8	322
124-47-0	217	130-15-4	355	137-26-8	208
124-58-3	517	130-22-3	499	137-30-4	554
124-63-0	180, 730, 744	130-24-5	905	137-32-6	182
124-64-1	534, 552	130-26-7	423	137-40-6	654
124-65-2	529	130-40-5	877	138-37-4	265
124-68-5	116	130-86-9	475	138-52-3	832
124-70-9	519	130-95-0	482, 712	138-84-1	645
124-73-2	139	131-08-8	529	138-89-6	323
124-76-5	250	131-11-3	324	139-02-6	653
124-83-4	225	131-14-6	307	139-13-9	190
124-87-8	255	131-18-0	308	139-59-3	369
125-46-2	902	131-22-6	370	139-65-1	308
125-58-6	883	131-28-2	357	139-66-2	331
126-00-1	286	131-48-6	857	139-85-5	377
126-07-8	443	131-56-6	317	140-10-3	303
126-11-4	216	131-69-1	376	140-11-4	279
126-14-7	833	131-74-8	627	140-22-7	329
126-17-0	855	131-91-9	364	140-28-3	309
126-18-1	854	132-27-4	653	140-29-4	280
126-29-4	841	132-28-5	551	140-40-9	397
126-30-7	152	132-29-6	502	140-66-9	365
126-33-0	485	132-53-6	364	140-87-4	136

140-88-5	158	146-77-0	866	191-24-2	276
140-89-6	524	147-14-8	633	192-97-2	277
140-93-2	532	147-71-7	203	195-19-7	276
140-95-4	151	147-73-9	203	195-84-6	492
141-22-0	201	147-82-0	390	198-55-0	367
141-28-6	112, 244	147-85-3	712, 794	206-44-0	334
141-32-2	123	147-93-3	385	208-96-8	258
141-43-5	115	147-94-4	871	213-14-3	376
141-52-6	651	148-24-3	449	217-59-4	394
141-53-7	651	148-82-3	792	218-01-9	303
141-75-3	129	149-29-1	468	222-93-5	309
141-78-6	157	149-30-4	457	229-87-8	470
141-79-7	179	149-32-6	822	230-27-3	409
141-82-2	178	149-44-0	531	244-69-9	419
141-84-4	484	149-45-1	509	248-13-5	408
141-86-6	427	149-74-6	517	252-47-1	275
141-90-2	491	149-91-7	336	253-52-1	472
141-97-9	158	150-13-0	262	253-66-7	423
142-04-1	268	150-25-4	784	253-82-7	482
142-08-5	448	150-30-1	793	254-60-4	465
142-09-6	171	150-39-0	172	254-79-5	465
142-18-5	168	150-46-9	539	259-56-3	276
142-29-0	238, 239	150-60-7	309	259-57-4	409
142-45-0	109	150-61-8	329	259-79-0	284
142-59-6	512	150-68-5	302	259-84-7	427
142-62-1	129	150-76-5	347	260-94-6	399
142-63-2	473	150-78-7	319	271-44-3	450
142-64-3	474	150-86-7	893	271-63-6	407
142-68-7	487	150-90-3	637	271-89-6	409
142-71-2	633	151-10-0	289, 319, 737	272-13-2	409
142-72-3	641	151-21-3	530	273-09-6	409
142-73-4	173, 744, 790	151-50-8	600	273-40-5	408
142-87-0	529	151-56-4	408	273-77-8	409
142-90-5	155	151-63-3	114	274-07-7	506
142-96-1	126	152-02-3	402	274-40-8	451
143-07-7	178	152-93-2	834	274-76-0	408
143-18-0	646	153-18-4	905	275-51-4	271
143-19-1	614, 653	153-78-6	264	275-94-5	427
143-24-8	206	153-98-0	897	280-57-9	427
143-27-1	131	154-08-5	786	280-64-8	503
143-28-2	193	154-17-6	820	281-23-2	220
143-62-4	849	154-41-6	889	286-20-4	232
143-66-8	532	154-42-7	490	286-62-4	234
143-67-9	902	156-06-9	375, 653	287-23-0	227
144-02-5	650	156-10-5	364	287-24-6	843
144-48-9	173	156-38-7	341	287-27-4	489
144-55-8	609	156-39-8	342	287-92-3	235
144-62-7	193	156-54-7	650	288-13-1	476
144-67-2	840	156-57-0	783	288-32-4	450
144-68-3	841	156-59-2	141	288-88-0	492
144-80-9	261	156-60-5	141	288-94-8	488
144-83-2	485	156-73-5	879	290-37-9	476
144-90-1	777	156-83-2	420	291-21-4	495
145-42-6	532	157-03-9	784	291-64-5	229
146-14-5	877	176-32-9	241	294-93-9	633
146-48-5	498	188-87-4	396	295-17-0	241
146-68-9	452	191-07-1	304, 938	295-37-4	486

297-90-5	447	328-42-7	193	370-69-4	553
298-07-7	502	328-50-7	193	372-09-8	136
298-14-6	599	329-56-6	889	372-31-6	164, 213
298-46-4	418	329-71-5	327	372-75-8	781
298-81-7	906	329-98-6	374	372-97-4	876
298-93-1	489	330-54-1	315	373-44-4	137
298-96-4	494	331-39-5	294	374-12-9	256
299-11-6	460	333-18-6	161	375-22-4	195
299-28-5	631	333-20-0	603	375-96-2	196
299-42-3	875	333-27-7	189	378-44-9	843
300-39-0	785	333-49-3	489	379-68-0	852
301-04-2	640	333-93-7	137	379-79-3	438
302-01-2	580	334-48-5	129, 178	381-73-7	146
302-17-0	131	334-50-9	202	382-45-6	842
302-27-2	399	334-88-3	137	383-63-1	164, 169
302-66-9	188	335-05-7	730	383-64-2	164
302-72-7	774	335-27-3	196, 254	389-08-2	886
302-79-4	903	335-57-9	195	392-12-1	452
302-95-4	651	335-67-1	194	392-56-3	337
303-04-8	213	338-69-2	774	394-31-0	339
303-07-1	317	338-83-0	196	401-56-9	159
303-21-9	327	339-72-0	782	404-86-4	295
303-43-5	847	340-05-6	376	405-50-5	336
303-81-1	890	340-06-7	375	405-71-2	712
303-97-9	868	341-92-4	335	405-79-8	336
303-98-0	868	343-27-1	444	407-24-9	119
304-06-3	375	343-65-7	454	407-25-0	213
304-20-1	445	343-94-2	495	407-41-0	893
304-21-2	443	345-78-8	377	408-35-5	653
304-55-2	147	350-03-8	398	420-04-2	574
304-59-6	646	350-46-9	335	420-37-1	658
304-88-1	279	352-21-6	777	420-56-4	513
304-91-6	344	352-32-9	336	421-06-7	121
305-01-1	438	352-70-5	336	421-50-1	169, 213
305-03-3	865	352-93-2	146	422-05-9	195
305-33-9	453	352-97-6	787	425-75-2	164
305-84-0	781	355-02-2	254	429-06-1	206, 534
306-07-0	366	355-42-0	196	429-07-2	205, 533
306-67-2	202	355-68-0	254	429-41-4	699, 700, 720
309-43-3	484	357-57-3	865	429-42-5	204, 533
309-88-6	145	358-23-6	214, 329	431-03-8	118
311-03-5	552	358-71-4	882	431-06-1	140
311-28-4	203	359-05-7	132, 605	432-68-8	839
311-89-7	196	360-65-6	851	434-13-9	852
312-30-1	285	360-68-9	846	434-16-2	845
313-04-2	845	361-37-6	370	437-11-6	535
315-22-0	869	363-03-1	371	438-08-4	850
316-42-7	437	363-72-4	366	438-22-2	842
316-46-1	877	366-18-7	412, 723	438-23-3	842
316-66-5	447	366-29-0	384	440-17-5	492
319-84-6	249	367-12-4	336	441-38-3	276
326-91-0	488	367-51-1	532	442-16-0	426
327-56-0	793	367-57-7	214	442-51-3	444
327-57-1	793	368-39-8	213	443-72-1	458
327-62-8	646	369-07-3	829	445-29-4	335
327-97-9	298	369-35-7	335	445-38-9	335

446-72-0	465	475-31-0	851	490-55-1	426
447-31-4	306	475-38-7	354	490-79-9	317
448-37-3	336	476-28-8	456	490-83-5	820
448-71-5	401	476-66-4	332	491-33-8	457
451-13-8	318	477-29-2	819	491-35-0	462
451-40-1	306	477-30-5	868	491-37-2	422
452-58-4	426	477-73-6	484	492-11-5	882
452-86-8	318	478-61-5	411	492-17-1	284
453-17-8	167	478-94-4	456	492-22-8	491
454-29-5	789	478-95-5	452	492-27-3	454
454-41-1	792	479-18-5	431	492-37-5	374
454-49-9	212	479-27-6	308	492-39-7	889
456-22-4	335	479-59-4	454	492-41-1	889
456-27-9	359	479-61-8	866	492-62-6	825
458-09-3	335	480-16-0	464	492-70-6	338
458-35-5	304	480-22-8	270	492-94-4	442
458-37-7	869	480-41-1	465	492-99-9	231
461-06-3	781	480-44-4	431	493-01-6	241
461-58-5	136	480-63-7	392	493-02-7	241
461-72-3	873	480-66-0	376	493-52-7	352
462-06-6	335	480-68-2	466	494-03-1	285
462-08-8	405	480-96-6	409	494-38-2	399
462-10-2	789	481-20-9	846	495-08-9	317
462-18-0	211	481-21-0	846	495-20-5	448
462-60-2	781	481-26-5	853	495-48-7	271
462-94-2	137	481-29-8	842	495-69-2	780
463-49-0	112	481-39-0	341	495-84-1	453
463-51-4	177, 459	481-42-5	340	496-11-7	343
463-58-1	568	481-74-3	318	496-16-2	430
463-82-1	189	482-05-3	328	496-41-3	409
464-10-8	120	482-74-6	424	496-67-3	121
464-43-7	224	482-89-3	450	496-78-6	393
464-48-2	225	483-04-5	401, 498	497-19-8	610
464-49-3	225	483-55-6	376	497-23-4	423
464-72-2	277	483-65-8	379	497-25-6	467
465-42-9	837	484-11-7	433	497-38-1	253
466-80-8	875	484-20-8	863	497-59-6	456
466-81-9	876	484-50-0	306	497-76-7	818
467-15-2	467	485-35-8	424	498-02-2	340
467-55-0	851	485-47-2	358	498-07-7	817
467-62-9	379	485-49-4	863	498-21-5	188
468-28-0	251	485-71-2	752, 867	498-23-7	134
469-59-0	454	485-89-2	447	498-24-8	179
470-38-2	838	486-25-9	335	498-59-9	785
470-82-6	248	486-35-1	431	498-66-8	253
471-46-5	193	486-84-0	443	498-94-2	794
472-61-7	836	487-21-8	475	499-06-9	322
472-70-8	839	487-51-4	226	499-83-2	436
472-86-6	904	487-79-6	454	500-05-0	423
472-92-4	838	487-90-1	894	500-22-1	478
472-93-5	838	488-23-3	384	500-38-9	365
473-30-3	268	488-43-7	824	500-40-3	337
473-34-7	312	488-82-4	818	500-99-2	319
474-07-7	414	488-93-7	441	501-30-4	454
474-25-9	844	489-84-9	337	501-52-0	375
474-62-4	844	490-02-8	407	501-53-1	282
474-67-9	843	490-11-9	479	501-58-6	318

501-65-5	328	512-35-6	274	524-36-7	896
501-97-3	342	512-47-0	131	524-38-9	342
501-98-4	339	512-56-1	542	524-42-5	355
502-41-0	229	512-69-6	831	524-63-0	424
502-42-1	229	513-35-9	116	524-64-4	308
502-49-8	234	513-36-0	174	524-80-1	418
502-65-8	839	513-38-2	174	524-97-0	476
502-72-7	234	513-44-0	187	525-03-1	264
502-74-2	177	513-48-4	173	525-05-3	520
503-17-3	129, 249	513-53-1	122	525-27-9	293
503-28-6	117	513-81-5	148	525-64-4	334
503-29-7	408	513-85-9	122	525-79-1	442
503-30-0	468	513-86-0	171	525-82-6	440
503-49-1	172	514-10-3	219	526-08-9	472
504-02-9	231	514-35-5	129	526-26-1	654
504-15-4	365	514-76-1	836	526-31-8	792
504-17-6	489	514-78-3	836, 837	526-55-6	495
504-20-1	196	514-85-2	904	526-75-0	323
504-24-5	405	515-30-0	270, 395	526-94-3	652
504-29-0	405	515-40-2	299	526-98-7	827
504-31-4	480	515-42-4	529	526-99-6	823
504-63-2	197	515-74-2	532	527-07-1	652, 825
505-22-6	434	515-83-3	131	527-17-3	332
505-23-7	437	515-84-4	163	527-52-6	822
505-48-6	202	516-05-2	185	527-53-7	384
505-54-4	169	516-12-1	173	527-54-8	393
505-66-8	445	516-88-7	846	527-60-6	393
505-95-3	171	517-09-9	849	527-61-7	322
506-03-6	131	517-10-2	847	527-62-8	263
506-12-7	167	517-23-7	109	527-72-0	491
506-17-2	217	517-28-2	880	527-76-4	268
506-30-9	117	517-51-1	385	527-85-5	386
506-46-7	168	518-05-8	354	528-29-0	325
506-59-2	148	518-28-5	893	528-44-9	392
506-64-9	607	518-82-1	332	528-45-0	325
506-65-0	579	519-05-1	365	528-71-2	447
506-68-3	574	519-23-3	437	528-76-7	325
506-78-5	574	519-44-8	327	528-93-2	386
506-96-7	108	519-62-0	866	529-19-1	389
507-02-8	110	519-73-3	394	529-23-7	262
507-19-7	120	519-88-0	311	529-34-0	383
507-20-0	125	520-26-3	444	529-36-2	385
507-25-5	130	520-33-2	444	529-44-2	886
507-28-8	535	520-36-5	492	529-59-9	824
507-32-4	513	520-85-4	852	529-64-6	395
507-47-1	131	521-11-9	852	529-69-1	453, 906
508-04-3	223	521-24-4	519	529-86-2	270
508-75-8	819	521-31-3	455	530-36-9	264
509-14-8	208	521-35-7	418	530-43-8	866
509-34-2	483	521-74-4	428	530-44-9	320
510-13-4	346	522-66-7	880	530-47-2	330
510-64-5	851	522-75-8	490	530-48-3	329
511-08-0	437	523-21-7	647	530-50-7	330
511-09-1	437	523-27-3	310	530-55-2	319
511-18-2	853	523-44-4	520	530-57-4	381
512-04-9	849	523-67-1	642	530-59-6	340

530-62-1	419	537-65-5	307	543-80-6	627
530-75-6	261	537-92-8	259	543-81-7	628
530-85-8	363	537-98-4	340	543-88-4	631
530-93-8	383	538-14-7	773	543-94-2	654
531-14-6	423	538-28-3	281	544-01-4	174
531-53-3	408	538-51-2	281	544-15-0	627
531-55-5	408	538-56-7	303	544-17-2	631
531-57-7	408	538-58-9	308, 674	544-25-2	229
531-72-6	505	538-62-5	329	544-40-1	140
531-81-7	423	538-74-9	309	544-47-8	298
531-84-0	337	538-75-0	246	544-62-7	117
531-85-1	275	538-81-8	328	544-63-8	189
531-91-9	328	539-03-7	296	544-76-3	168
531-99-7	260	539-12-8	377	544-92-3	573
532-02-5	531	539-15-1	338	545-06-2	210
532-27-4	296	539-17-3	262	545-26-6	851
532-32-1	650	539-30-0	281	546-06-5	423
532-40-1	900	539-35-5	900	546-18-9	845
532-54-7	344	539-47-9	442	546-56-5	520
532-55-8	278	539-64-0	785	546-67-6	641
532-78-5	419	539-88-8	162	546-68-9	655
532-82-1	303	540-04-5	840	546-74-7	531
533-01-7	633	540-05-6	840	546-98-4	844
533-30-2	402	540-23-8	388	547-25-1	834
533-48-2	872	540-37-4	343	547-57-9	530
533-58-4	344	540-38-5	344	547-58-0	517
533-62-0	790	540-54-5	200	548-04-9	342
533-67-5	820	540-59-0	141	548-24-3	637
533-73-3	275	540-67-0	163	548-42-5	401
533-75-5	395	540-69-2	627	548-43-6	437
533-87-9	112	540-72-7	616	548-62-9	305
533-90-4	632	540-80-7	128	548-65-2	366
534-00-9	120	540-84-1	215	548-93-6	339
534-07-6	140	540-88-5	123	549-38-2	350
534-15-6	104	541-14-0	781	549-56-4	482
534-17-8	569	541-15-1	781	550-33-4	895
534-22-5	459	541-41-3	159, 216	550-44-7	352
534-52-1	326	541-47-9	148	550-57-2	341
534-85-0	264	541-59-3	481	550-60-7	362
534-94-1	403	541-73-1	312	550-74-3	473
535-11-5	158	541-88-8	132	551-06-4	357
535-15-9	160	542-05-2	106	551-16-6	405
535-46-6	501	542-11-0	268	552-16-9	360
535-80-8	297	542-18-7	226	552-23-8	364
535-83-1	492	542-28-9	496	552-32-9	358
535-87-5	307	542-32-5	775	552-38-5	641
536-10-6	368	542-55-2	174	552-46-5	356
536-33-4	439	542-59-6	108	552-62-5	464
536-40-3	297	542-69-8	127	552-79-4	350
536-69-6	443	542-90-5	163	552-86-3	442
536-74-3	370	542-92-7	234, 635, 648	552-89-6	358
536-75-4	440	543-21-5	109	553-12-8	895
536-90-3	269	543-24-8	772	553-17-3	337
537-00-8	632	543-27-1	174	553-24-2	465
537-17-7	402	543-38-4	781	553-26-4	412
537-47-3	375	543-59-9	116	553-39-9	341
537-64-4	637	543-63-5	630	553-53-7	466

553-54-8	641	561-20-6	417	579-92-0	328
553-70-8	642	562-48-1	792	580-13-2	291
553-79-7	363	562-90-3	528	580-15-4	406
553-90-2	151	563-41-7	201	580-17-6	406
553-97-9	349	563-47-3	299	581-05-5	791
554-00-7	312	563-63-3	606, 648, 707, 937	581-40-8	323
554-01-8	458	563-68-8	657	581-42-0	323
554-05-2	661	563-79-1	149	581-64-6	490
554-12-1	188	563-80-4	183	581-89-5	361
554-13-2	586	563-83-7	174	581-96-4	356
554-68-7	212	564-02-3	215	582-25-2	645
554-70-1	539	564-36-3	437	582-52-5	821
554-73-4	553	565-00-4	225	582-60-5	431
554-84-7	362	565-48-0	255	583-39-1	457
554-91-6	824	565-50-4	255	583-59-5	252
554-95-0	275	565-80-0	147	583-61-9	455
555-06-6	650	566-57-5	854	583-75-5	263
555-16-8	358	566-58-5	854	583-78-8	314
555-21-5	360	566-78-9	842	583-93-7	784
555-30-6	785	568-93-4	358	584-02-1	195
555-31-7	626	569-61-9	379	584-04-3	147
555-44-2	215	571-61-9	323	584-08-7	599
555-45-3	215	571-92-6	850	584-13-4	407
555-48-6	787	572-09-8	816	584-48-5	290
555-68-0	361	573-56-8	327	584-84-9	389
555-75-9	626	573-57-9	317	584-85-0	779
556-10-5	363	573-58-0	304	585-07-9	128
556-32-1	642	573-83-1	646	585-47-7	274
556-33-2	784	573-97-7	291	585-48-8	428
556-50-3	787	574-15-2	275	585-76-2	287
556-52-5	443	574-16-3	275	585-79-5	291
556-56-9	114	574-25-4	883	585-84-2	110
556-63-8	641	574-66-3	277	585-99-9	828
556-65-0	587	574-93-6	644	586-11-8	327
556-67-2	520	574-98-1	290	586-38-9	347
556-88-7	190	575-44-0	318	586-75-4	288
556-89-8	190	575-47-3	528	586-76-5	288
556-91-2	626	576-15-8	398	586-77-6	290
556-99-0	217	576-19-2	864	586-78-7	291
557-01-7	448	576-24-9	314	586-89-0	260
557-09-5	661	576-26-1	323	586-96-9	363
557-11-9	114	576-36-3	823	587-04-2	297
557-17-5	188	577-11-7	530	587-85-9	636
557-20-0	637	577-19-5	291	587-90-6	326
557-21-1	624	577-56-0	260	587-98-4	516
557-24-4	178	577-71-9	327	589-02-6	502
557-41-5	661	577-85-5	446	589-15-1	288
557-59-5	204	578-57-4	287	589-16-2	333
557-61-9	191	578-66-5	406	589-17-3	288
557-66-4	158	578-68-7	406	589-18-4	389
557-99-3	110	578-76-7	459	589-21-9	292
558-13-4	130	578-86-9	431	589-34-4	185
558-17-8	127	578-95-0	400	589-90-2	247
558-32-7	207, 535	579-10-2	348	589-92-4	252
560-09-8	225	579-18-0	278	589-93-5	456
561-07-9	426	579-43-1	338	590-18-1	123

590-28-3	600	598-50-5	189	607-93-2	390
590-46-5	780	598-53-8	177	608-33-3	310
590-47-6	780	598-55-0	183	608-36-6	139
590-73-8	151	598-58-3	186	608-37-7	138
590-92-1	120	598-75-4	183	608-40-2	152
591-08-2	110	598-78-7	134	608-66-2	822
591-22-0	456	599-04-2	881	608-71-9	366
591-23-1	252	599-66-6	332	608-91-3	391
591-27-5	267	599-67-7	329	608-92-4	391
591-31-1	268	600-37-3	172	609-19-8	391
591-35-5	315	601-89-8	363	609-67-6	344
591-49-1	252	602-01-7	327	609-72-3	324
591-50-4	343	602-09-5	283, 712	609-89-2	314
591-54-8	406	602-38-0	326	609-93-8	326
591-55-9	406	602-55-1	370	609-99-4	327
591-76-4	185	602-60-8	358	610-02-6	392
591-80-0	113	602-94-8	366	610-09-3	230
591-87-7	113	602-99-3	394	610-27-5	363
591-93-5	194	603-00-9	448	610-28-6	325
592-04-1	589	603-02-1	327	610-30-3	325
592-13-2	151	603-11-2	363	610-39-9	327
592-20-1	107	603-12-3	325	610-54-8	326
592-31-4	189	603-32-7	549	610-71-9	310
592-41-6	170	603-33-8	659	610-72-0	322
592-42-7	169	603-34-9	394	610-94-6	349
592-57-4	229	603-35-0	549	611-01-8	322
592-76-7	168	603-36-1	659	611-08-5	467
592-84-7	126	603-45-2	379	611-10-9	244, 247, 248
592-85-8	589	603-54-3	331	611-32-5	462
593-45-3	191	603-71-4	361	611-33-6	422
593-51-1	182	603-79-2	322	611-34-7	406
593-56-6	181	603-83-8	351	611-35-8	422
593-81-7	215	604-35-3	847	611-71-2	346
593-85-1	167	604-44-4	300	611-73-4	278
593-95-3	568	604-53-5	284	611-95-0	278
594-09-2	542	604-59-1	465	611-97-2	322
594-15-0	210	604-68-2	825	611-98-3	307
594-18-3	138	604-69-3	826	611-99-4	317
594-20-7	141	604-88-6	337	612-14-6	273
594-27-4	694, 656, 658	605-48-1	312	612-37-3	464
594-36-5	116	605-60-7	364	612-42-0	794
594-44-5	155	605-69-6	326	612-60-2	462
594-61-6	172	605-91-4	473	612-62-4	422
594-65-0	210	605-94-7	868	612-71-5	394
595-30-2	225	606-20-2	327	612-94-2	374
595-40-4	791	606-22-4	325	612-95-3	472
595-46-0	151	606-23-5	342	612-98-6	364
595-49-3	153	606-26-8	359, 360	613-12-7	349
595-90-4	656	606-35-9	393	613-31-0	317
596-03-2	428	606-40-6	300	613-78-5	357
596-38-3	472	606-68-8	888	614-75-5	341
597-12-6	828	606-81-5	326	614-80-2	258
597-43-3	152	606-83-7	331	615-18-9	419
597-71-7	831	607-24-9	362	615-20-3	419
597-80-8	502	607-34-1	467	615-28-1	372
598-04-9	140	607-57-8	361	615-43-0	343
598-31-2	119	607-92-1	322, 324	615-57-6	310

615-58-7	310	621-59-0	345	626-97-1	218
615-67-8	299	621-66-9	340	627-06-5	189
615-93-0	313	621-82-9	303	627-11-2	132
615-95-2	432	621-84-1	280	627-20-3	195
616-02-4	134	622-15-1	330	627-21-4	195
616-03-5	884	622-21-9	315	627-45-2	166
616-04-6	880, 881, 884	622-32-2	272	627-83-8	161
616-06-8	793	622-47-9	352	627-84-9	161
616-13-7	116	622-51-5	390	627-93-0	148
616-29-5	137	622-78-6	281	627-95-2	779
616-38-6	149	622-97-9	353	628-02-4	129
616-44-4	463	623-05-2	339	628-13-7	479
616-45-5	716	623-12-1	297	628-28-4	128
616-47-7	459	623-24-5	310	628-41-1	229
616-75-1	281	623-25-6	315	628-63-7	116
617-33-4	160	623-26-7	316	628-73-9	129
617-35-6	163	623-27-8	381	628-81-9	126
617-45-8	780	623-33-6	787	629-04-9	168
617-48-1	178	623-42-7	183	629-05-0	193
617-52-7	151	623-49-4	160	629-11-8	170
617-73-2	172	623-51-8	163	629-14-1	162
617-84-5	145	623-65-4	193	629-20-9	234
617-86-7	540	623-70-1	159	629-25-4	651
617-89-0	442	623-73-4	160	629-31-2	168
618-36-0	349	623-84-7	200	629-41-4	192
618-39-3	272	623-91-6	145	629-45-8	126
618-41-7	274	624-03-3	161	629-54-9	131
618-51-9	344	624-31-7	344	629-59-4	205
618-80-4	314	624-38-4	318	629-70-9	130
618-95-1	352	624-49-7	145	629-96-9	117
619-04-5	322	624-60-2	184	629-97-0	154
619-05-6	307	624-64-6	123	630-01-3	168
619-15-8	327	624-65-7	198	630-02-4	191
619-17-0	266	624-73-7	146	630-08-0	568
619-19-1	363	624-84-0	166	630-10-4	528
619-21-6	295	624-91-9	186	630-18-2	125
619-23-8	360	624-92-0	150	630-60-4	853
619-31-8	323	625-35-4	135	631-36-7	534
619-42-1	349	625-45-6	181	631-40-3	209, 754
619-50-1	352	625-47-8	133	631-61-8	627
619-58-9	344	625-51-4	104	632-12-2	791
619-65-8	306, 778	625-52-5	189	632-22-4	208
619-66-9	295, 778	625-53-6	163	632-25-7	381
619-67-0	338	625-77-4	136	632-46-2	322
619-73-8	360	626-05-1	428	632-51-9	385
619-84-1	320	626-15-3	310	632-52-0	385
620-22-4	389	626-35-7	163	632-56-4	206
620-24-6	339	626-39-1	390	632-69-9	528
620-40-6	390	626-48-2	463	632-92-8	394
620-45-1	509	626-51-7	184	632-93-9	428
620-50-8	322	626-56-2	460	632-99-5	379
621-02-3	332	626-58-4	460	632-99-2	443
621-09-0	328	626-60-8	421	633-03-4	286
621-23-8	392	626-61-9	421	633-65-8	411
621-37-4	341	626-64-2	448	633-71-6	340
621-42-1	258	626-72-2	789	633-96-5	521

633-99-8	300	646-31-1	204	700-57-2	221
634-35-5	483	652-04-0	349	700-58-3	221
634-36-6	392	653-89-4	884	701-99-5	369
634-65-1	324	657-26-1	791	703-59-3	338
634-66-2	382	657-27-2	791	703-80-0	398
634-67-3	391	657-84-1	533	704-00-7	306
634-83-3	382	658-78-6	363	704-15-4	788
634-90-2	382	660-68-4	142	706-35-2	467
634-93-5	391	660-88-8	779	708-06-5	340
635-21-1	263	661-19-8	154	708-79-2	463
635-51-8	375	661-69-8	639	709-49-9	344
635-65-4	412	665-66-7	221	709-63-7	391
635-78-9	647	666-99-9	112	712-74-3	383
635-93-8	302	670-54-2	205	716-53-0	297
636-04-4	385	672-15-1	789	719-22-2	311
636-09-0	316	673-40-5	287	723-62-6	270
636-30-6	391	673-49-4	857	725-12-2	435
636-70-4	212	673-84-7	113	732-26-3	390
636-73-7	479	674-26-0	885	733-52-4	171
636-88-4	889	674-76-0	187	734-77-0	284
636-98-6	344	674-82-8	459, 493	741-03-7	206
637-01-4	384	675-20-7	413, 496	747-90-0	845
637-39-8	211	676-97-1	517	751-38-2	385
637-44-5	374	677-53-6	385	754-05-2	548
637-59-2	375	677-69-0	167	754-05-2	750
637-84-3	797	680-31-9	433, 514, 721	754-34-7	196
637-87-6	299	681-84-5	534	758-96-3	152
637-88-7	231	684-16-2	169	759-24-0	143
637-92-3	126	684-19-5	734	760-21-4	159
638-32-4	202	685-73-4	824	760-32-7	509
638-38-0	642	685-87-0	142	760-93-0	179
638-39-1	658	685-91-6	142	761-65-9	139
638-53-9	211	687-38-7	788	762-21-0	142
638-95-9	223	688-71-1	551	762-62-9	152
639-58-7	659	688-73-3	658	762-72-1	499
639-81-6	438	688-74-4	537	763-10-0	879
640-19-7	164	690-02-8	152	763-29-1	187
640-60-8	375	691-38-3	187	764-13-6	151
641-13-4	270	693-02-7	171	764-35-2	171
641-36-1	861	693-23-2	154	764-85-2	194
641-70-3	363	693-65-2	117	766-39-2	151
641-85-0	853	693-72-1	217	767-00-0	306
642-31-9	270	693-98-1	459	767-15-7	403
642-44-4	402	694-53-1	522	767-92-0	425
642-71-7	392	694-59-7	479	768-32-1	541
642-78-4	867	694-80-4	289, 290	768-33-2	510
643-13-0	878	694-80-4	737	768-90-1	221
643-43-6	327	695-34-1	405	768-91-2	252
644-32-6	278	696-54-8	478	768-92-3	222
644-76-8	817	696-62-8	343	768-93-4	222
644-97-3	522	697-82-5	393	768-94-5	221
645-48-7	375	697-91-6	313	768-95-6	221
645-49-8	380	698-69-1	320	769-42-6	431
645-96-5	501	700-12-9	367	769-90-1	222
646-04-8	195	700-13-0	393	770-12-7	522
646-06-0	435	700-16-3	469	770-71-8	250
646-30-0	191	700-56-1	252	771-29-9	383

771-51-7	451	828-60-3	448	877-24-7	645
771-61-9	366	828-73-9	716	877-44-1	391
771-97-1	308	829-84-5	737	879-39-0	382
772-14-5	371	829-85-6	536, 741	880-52-4	221
772-15-6	371	830-03-5	362	882-06-4	361
772-33-8	341	830-09-1	347	882-33-7	329
773-76-2	429	830-81-9	356	883-99-8	350
775-12-2	511	830-96-6	451	886-38-4	329
776-34-1	266	831-23-2	385	890-38-0	872
779-02-2	349	831-59-4	273	892-21-7	361
781-43-1	322	831-61-8	333	893-33-4	356
781-73-7	260	835-64-3	447	901-93-9	850
782-92-3	279	836-30-6	361	902-47-6	326
784-04-3	259	838-85-7	511	914-45-4	204
784-62-3	490	840-78-8	413	915-35-5	851
787-70-2	284	842-07-9	273	917-23-7	488
789-25-3	551	845-10-3	651	917-61-3	611
791-28-6	550	845-46-5	651	919-62-0	540
791-31-1	551	846-70-8	519	920-46-7	179
793-41-9	268	847-25-6	901	920-66-1	170
804-63-7	482	852-38-0	412	921-01-7	783
811-93-8	187	853-35-0	528	921-53-9	646
813-19-4	639	853-68-9	500	922-32-7	782
814-68-6	111	860-22-0	515	922-67-8	188
814-80-2	632	863-57-0	652	923-73-9	118
814-81-3	633	865-34-9	641	924-49-0	514
814-95-9	654	865-44-1	585	924-49-2	777
818-08-6	636	865-47-4	645	926-39-6	115
818-88-2	201	865-48-5	650	928-04-1	626
819-19-2	508	866-83-1	645	928-49-4	171
819-83-0	512	866-97-7	203, 754	929-75-9	137
821-06-7	138	867-13-0	540	930-28-9	227, 241
821-48-7	118	867-81-2	891	930-46-1	239
822-16-2	654	868-14-4	645	930-88-1	185
822-23-1	191	868-19-9	654	931-19-1	472
822-36-6	459	868-77-9	172	931-20-4	496
822-67-3	232	869-24-9	134	931-71-5	231
822-87-7	227, 236	869-29-4	111	931-87-3	234
823-18-7	231	870-46-2	125	931-88-4	234
823-76-7	219	870-50-8	139	932-90-1	272
823-96-1	525, 541	870-93-9	789	933-18-6	538
823-96-1	717, 718	872-35-5	457	933-52-8	256
824-35-1	632	872-36-6	497	934-34-9	446
824-72-6	522	872-50-4	462	935-30-8	418
824-79-3	532	872-53-7	235	935-31-9	228
824-94-2	347	872-85-5	478	935-56-8	222
825-90-1	532	872-93-5	462	935-92-2	392
826-73-3	277	873-32-5	298	935-95-5	382
826-74-4	396	873-55-2	529	937-14-4	301
826-81-3	447	873-66-5	353	938-55-6	431
827-19-0	530	873-69-8	478	938-79-4	371
827-21-4	530	873-74-5	263	938-85-2	459
827-52-1	306	873-83-6	407	939-26-4	290
827-54-3	396	874-14-6	433	939-52-6	301
828-01-3	342	874-23-7	219	940-31-8	369
828-51-3	220	874-42-0	312	940-69-2	882

940-71-6	596	1010-48-6	300	1116-76-3	215
941-75-3	237	1010-95-3	463	1116-98-9	125
941-98-0	260	1012-72-2	311	1117-86-8	192
943-89-5	347	1020-31-1	311	1118-02-1	545
946-82-7	232	1024-41-5	283	1118-12-3	128
947-04-6	407	1024-99-3	882	1119-34-2	779
947-42-2	511	1025-70-4	617	1119-94-4	155
947-84-2	284	1032-65-1	871	1119-97-7	205, 754
948-60-7	895	1034-39-5	549	1120-01-0	531
951-55-3	792	1041-01-6	785	1120-02-1	754
951-77-9	871	1045-69-8	856	1120-44-1	633
951-78-0	872	1046-56-6	480	1120-49-6	141
952-23-8	426	1048-08-4	536	1121-22-8	242
957-75-5	865	1059-86-5	845	1121-60-4	478
958-80-5	320	1066-26-8	109, 650	1121-89-7	474
959-10-4	284	1066-45-1	659, 694	1122-17-4	141
959-28-4	309	1066-54-2	513	1122-58-3	329
960-71-4	549	1067-12-5	534, 552	1122-58-3	413, 718
961-07-9	872	1068-57-1	105	1123-40-6	433
961-68-2	326	1068-90-2	142	1123-54-2	407
963-16-6	745	1069-66-5	200	1124-13-6	227
971-74-4	484	1069-79-0	874	1126-34-7	531
975-17-7	373	1070-89-9	529, 550	1126-71-2	323
977-96-8	430	1071-23-4	892	1126-79-0	294
979-92-0	860	1072-62-4	755	1128-87-6	397
981-40-8	505	1072-72-6	487	1132-21-4	319
989-38-8	483	1072-86-2	230	1132-68-9	786
992-59-6	502	1072-98-6	403	1134-35-6	432
992-92-7	658	1074-82-4	646	1135-24-6	340
993-13-5	517	1076-22-8	464	1135-66-6	251
993-50-0	659	1076-38-6	446	1136-89-6	887
994-32-1	658	1076-43-3	273	1137-41-3	263
994-89-8	637	1077-28-7	883	1137-42-4	279
995-45-9	537	1079-66-9	506, 700	1139-30-6	226
996-50-9	509	1079-66-9	741, 749, 751	1141-59-9	480
996-70-3	207	1079-71-6	365	1143-74-4	279
998-07-2	874	1081-34-1	485	1148-79-4	485
998-30-1	539	1083-48-3	466	1152-14-3	896
998-40-3	537, 754	1085-92-3	508	1152-90-5	478
999-90-6	128	1086-80-2	455	1153-05-5	549
999-97-3	514, 730	1088-56-8	883	1155-00-6	286
1000-84-6	157	1094-61-7	888	1163-67-3	465
1001-53-2	109	1095-90-5	884	1164-95-0	842
1002-62-6	650	1099-45-2	513	1170-02-1	785
1002-84-2	194	1100-88-5	502	1173-26-8	848
1003-32-3	489	1101-41-3	536	1173-88-2	890
1003-67-4	473	1107-26-2	836	1174-69-2	854
1003-73-2	473	1109-11-1	836	1176-74-5	382
1003-78-7	152, 433	1111-39-3	816	1177-87-3	849
1005-24-9	797	1112-67-0	203	1180-95-6	855
1005-56-7	371	1113-41-3	891	1185-53-1	216
1006-94-6	458	1113-59-3	120	1185-55-3	518
1007-32-5	281	1114-34-7	828	1186-49-8	652
1007-42-7	788	1114-41-6	885	1187-58-2	188
1008-72-6	500	1114-81-4	893	1187-84-4	792
1009-61-6	307	1115-84-0	905	1188-37-0	772
1009-93-4	514	1116-54-7	190	1188-58-5	153

1188-92-7	540	1298-55-1	660	1445-73-4	460
1188-95-0	215, 755	1301-96-8	608	1446-61-3	242
1189-94-2	885	1302-78-9	809	1450-14-2	514
1191-50-0	532	1303-58-8	579	1453-82-3	452
1191-95-3	228	1303-82-6	564	1454-85-9	167
1192-28-5	241	1303-96-4	609	1455-77-2	427
1192-40-1	507	1304-56-9	564	1457-42-7	477
1192-52-5	429	1305-62-0	567	1460-02-2	390
1193-92-6	478	1309-64-4	560, 933	1460-57-7	230
1194-65-6	313	1310-53-8	577	1461-15-0	631
1195-16-0	773	1310-58-3	601	1461-22-9	648, 658
1195-45-5	336	1310-65-2	587	1461-25-2	655
1197-17-7	776	1310-66-3	587	1461-94-5	227
1197-18-8	776	1310-73-2	612	1461-96-7	236
1197-55-3	267	1311-10-0	617	1461-97-8	237
1198-27-2	266	1312-43-2	584	1464-42-2	795
1198-33-0	464	1313-27-5	591	1464-44-4	817, 831
1199-03-7	483	1313-82-2	615	1465-25-4	357
1200-93-7	457	1313-84-4	615	1466-76-8	319
1202-34-2	436	1313-97-9	592	1468-83-3	399
1204-06-4	451	1314-28-9	668	1470-61-7	528
1205-72-7	403	1314-32-5	620	1470-94-6	340
1207-72-3	460	1314-56-3	597	1474-52-8	850
1210-12-4	305	1314-68-7	596, 669	1474-78-8	540
1212-08-4	370	1314-80-3	597	1476-23-9	114
1214-39-7	410	1314-85-8	598	1476-53-5	890
1215-59-4	411	1317-36-8	586	1481-70-5	797
1217-45-4	316	1321-74-0	332	1482-97-9	784
1218-34-4	774	1324-63-6	834	1484-09-9	453
1220-94-6	265	1327-53-3	561	1484-12-4	458
1224-69-7	413	1330-20-7	396	1484-13-5	497
1226-46-6	385	1330-43-4	609	1484-28-2	416
1235-21-8	499	1330-78-5	553	1490-04-6	760
1239-31-2	842	1332-52-1	563, 564, 628	1492-08-6	885
1239-45-8	53, 876	1333-74-0	581	1492-24-6	775
1245-13-2	412	1333-82-0	571	1493-13-6	189
1260-17-9	865	1333-86-4	130, 924	1498-88-0	493
1263-89-4	891	1344-09-8	615	1499-10-1	328
1264-51-3	816	1344-28-1	556	1499-17-8	521
1271-03-0	675, 721	1344-58-7	622	1501-84-4	222, 223
1271-19-8	658	1393-62-0	798, 799	1502-03-0	228
1271-24-5	633, 673	1395-18-2	271	1502-06-3	228
1271-28-9	644	1400-61-9	890	1503-53-3	260
1271-42-7	638, 741	1405-69-2	800	1515-76-0	107
1271-51-8	660	1405-87-4	862	1522-22-1	169, 639, 691
1271-55-2	626, 638	1414-45-5	793	1523-11-1	356
1277-43-6	673	1415-73-2	817	1528-74-1	326
1277-47-0	673	1429-59-0	153, 874	1530-32-1	513
1277-49-2	638, 738	1432-43-5	468	1530-45-6	513
1284-72-6	673	1435-55-8	880	1544-68-9	336
1287-13-4	647	1435-88-7	340	1552-12-1	234, 240
1291-32-3	711	1436-34-6	417	1552-12-1	234, 629, 634
1291-48-1	638	1436-59-5	242	1552-96-1	321
1293-87-4	638	1443-80-7	260	1559-02-0	247
1294-74-2	640	1445-23-4	293, 735	1562-94-3	271
1295-35-8	696	1445-69-8	472	1563-91-3	246

1564-64-3	287	1710-98-1	293	1923-70-2	204
1565-80-6	182	1711-02-0	344	1928-38-7	350
1569-69-3	233	1711-10-0	344	1936-15-8	521
1570-64-5	299	1720-32-7	330	1937-54-8	202
1570-65-6	314	1722-12-9	422	1941-24-8	207
1571-33-1	522	1722-26-5	504	1941-27-1	204
1574-41-0	194	1723-00-8	473, 794	1941-30-6	208, 754
1576-35-8	387	1730-91-2	183	1945-77-3	518
1591-31-7	344	1733-12-6	305	1948-33-0	294
1592-20-7	395	1739-84-0	433	1949-78-6	828
1600-27-7	643	1746-03-8	554	1949-89-9	820
1600-44-8	487	1750-12-5	403	1950-85-2	180
1602-97-7	131	1754-47-8	510	1972-28-7	142
1603-40-3	404	1755-15-3	219	1983-26-2	506
1603-41-4	405	1758-73-2	165	1985-12-2	292
1606-67-3	267	1762-95-4	560	1988-89-2	373
1608-26-0	514, 547, 548	1763-10-6	890	1989-52-2	847
1609-47-8	146	1772-03-8	823	1990-29-0	817
1609-86-5	127	1779-48-2	522	2001-45-8	536
1611-35-4	498	1779-49-3	519	2001-95-8	902
1615-02-7	298	1779-58-4	517	2002-24-6	115
1618-26-4	165	1783-81-9	350	2004-06-0	866
1621-91-6	477	1787-61-7	512	2004-14-0	516
1632-73-1	248	1798-04-5	294	2004-70-8	194
1633-22-3	365	1798-09-0	347	2009-64-5	887
1634-04-4	128	1805-32-9	314	2018-61-3	773
1635-31-0	53	1806-29-7	317	2026-48-4	744, 759
1637-39-4	906	1811-28-5	426	2026-48-4	760, 778
1641-41-1	340	1817-74-9	286	2030-63-9	423
1643-19-2	203	1820-81-1	422	2031-12-9	793
1643-20-5	154	1821-12-1	371	2031-67-6	518
1653-19-6	140	1822-51-1	420	2033-24-1	175, 432, 493
1655-07-8	247	1824-81-3	405	2035-66-7	595
1655-29-4	512	1826-28-4	472	2039-82-9	292
1656-44-6	325	1829-28-3	343	2042-37-7	287
1660-04-4	222	1832-07-1	596	2043-06-3	412
1663-39-4	124	1833-53-0	527	2044-56-6	516
1663-45-2	674	1838-08-0	191	2045-00-3	500
1663-45-2	726, 752	1846-92-2	332	2050-76-2	314
1666-13-3	510	1850-14-2	208	2051-28-7	425
1667-99-8	507	1863-63-4	627	2051-85-6	317
1668-10-6	786	1866-15-5	858	2052-07-5	288, 722
1670-46-8	219	1866-16-6	865	2057-49-0	471, 747
1674-10-8	247	1871-22-3	413	2057-84-3	218
1676-73-9	786	1871-96-1	201	2057-87-6	241
1678-91-7	247	1873-77-4	553	2058-46-0	899
1689-82-3	370	1875-48-5	267	2067-33-6	121
1694-31-1	123	1878-65-5	301	2068-78-2	902
1694-92-4	359	1878-66-6	301	2082-84-0	136, 753
1696-20-4	398	1878-68-8	291	2086-83-1	411
1699-51-0	455	1878-87-1	362	2090-05-3	631
1701-93-5	608	1878-91-7	291	2091-46-5	526
1707-03-5	511	1888-91-1	397	2092-17-3	563
1707-75-1	330	1910-42-5	432	2092-55-9	380, 519
1707-77-3	822	1913-93-5	386	2092-55-9	523, 530
1709-54-2	283	1918-77-0	490	2094-72-6	220
1709-66-9	536	1921-70-6	208	2094-98-6	223

2107-69-9	319	2321-07-5	53, 440	2498-66-0	272
2113-57-7	288	2338-20-7	392	2504-11-2	772
2116-65-6	411	2340-34-4	917	2507-55-3	173
2121-67-7	150	2344-80-1	543, 545	2508-19-2	393
2122-61-4	263	2348-71-2	489	2511-17-3	535
2127-03-9	436	2361-24-2	840	2516-34-9	228
2131-55-7	302	2361-96-8	773	2524-64-3	511
2136-75-6	550	2370-18-5	336	2530-10-1	397
2138-24-1	207, 754	2370-63-0	157	2533-00-8	467
2140-46-7	851	2370-88-9	535	2537-36-2	207
2150-48-3	647	2373-80-0	350	2538-85-4	513
2152-56-9	818	2379-55-7	433	2549-67-9	439
2153-98-2	889	2385-77-5	135	2552-55-8	790
2155-71-7	311	2386-53-0	530	2554-06-5	535
2158-04-5	368	2386-60-9	122	2563-97-5	390
2161-31-1	471	2387-71-5	779	2564-83-2	487
2170-03-8	177	2390-54-7	489	2566-90-7	154
2173-69-5	300	2392-39-4	849	2567-83-1	205
2176-62-7	469	2396-60-3	346	2570-63-0	657
2179-57-9	114	2402-78-0	429	2572-40-9	741
2181-42-2	215	2403-88-5	449	2578-28-1	795
2183-17-7	519, 887	2408-37-9	256	2578-45-2	420
2197-63-9	509	2411-89-4	304	2578-57-6	795
2208-07-3	157	2412-14-8	491	2591-17-5	883
2210-25-5	176	2414-98-4	642	2592-95-2	446
2213-63-0	430	2421-28-5	277	2605-67-6	516, 550
2216-12-8	361	2421-28-5	937	2610-05-1	506
2216-33-3	186	2423-65-6	476	2611-82-7	519
2216-51-5	251, 760	2425-51-6	515	2615-25-0	243
2216-69-5	351	2425-74-3	166	2620-63-5	772
2217-07-4	331	2425-95-8	412	2622-14-2	539
2218-94-2	466	2430-27-5	200	2622-26-6	470
2219-31-0	781	2433-84-9	476	2623-36-1	531, 652
2225-23-2	333	2435-53-2	382	2627-86-3	349
2227-29-4	510	2436-90-0	135	2627-95-4	512
2234-75-7	256	2436-96-6	326	2635-26-9	316
2235-00-9	497	2437-29-8	346	2636-31-9	345
2235-54-3	499	2439-54-5	192	2637-34-5	479
2235-54-3	627	2439-55-6	191	2642-52-6	463
2236-60-4	475, 476	2440-42-8	637	2644-70-4	580
2240-25-7	414	2444-68-0	395	2645-22-9	436
2243-27-8	194	2446-83-5	147	2646-71-1	888
2243-42-7	369	2451-01-6	255	2650-53-5	171, 754
2243-62-1	308	2457-47-8	429	2658-24-4	432
2243-76-7	401	2457-80-9	872	2665-04-5	845
2243-82-5	354	2465-27-2	53, 271	2679-01-8	458
2244-11-3	402	2466-76-4	398	2679-14-3	774
2244-40-8	530	2475-44-7	286	2687-43-6	281,
2246-46-0	489	2475-93-6	196	2688-77-9	455
2247-00-9	630	2482-00-0	112	2696-92-6	594
2290-65-5	545	2487-90-3	541	2700-22-3	281
2292-79-7	242	2491-20-5	774	2712-78-9	286
2294-76-0	407	2492-02-9	292	2719-27-9	230
2305-32-0	230	2492-87-7	830	2739-12-0	348
2305-34-2	211, 755	2494-56-6	865	2743-38-6	138
2309-49-1	487	2498-20-6	203, 754	2744-08-3	425

2744-09-4	425	3025-95-4	772	3383-21-9	311
2749-11-3	774	3029-19-4	378	3385-94-2	503
2757-04-2	372	3029-30-9	316	3387-26-6	441
2763-96-4	886	3029-40-1	330	3399-67-5	115
2765-29-9	228	3034-55-7	416	3400-45-1	235
2767-80-8	534, 552, 935	3051-09-0	465	3411-95-8	447
2767-84-2	225	3054-95-3	111	3413-58-9	439
2768-90-3	377, 473	3068-32-4	816	3416-05-5	872
2769-71-3	324	3076-34-4	615	3416-24-8	825
2769-72-4	127	3076-91-3	643	3425-46-5	603
2782-07-2	823	3084-10-4	536	3434-88-6	850
2797-68-4	874	3087-16-9	515	3443-45-6	378
2799-21-5	449	3097-08-3	516	3458-28-4	828
2818-58-8	831	3102-56-5	873	3460-18-2	310
2835-68-9	262	3105-95-1	473, 794	3463-36-3	361
2835-81-6	775	3109-63-5	203	3465-72-3	496
2835-82-7	776	3112-31-0	477	3468-11-9	431
2835-98-5	264	3118-85-2	824	3473-63-0	165
2835-99-6	264	3140-93-0	416	3481-20-7	382
2836-32-0	652	3141-04-6	233	3530-79-8	781
2840-02-0	263	3141-26-2	428	3535-83-9	207, 754
2840-28-0	263	3143-02-0	460	3562-69-4	493
2851-68-5	478	3144-16-9	226	3564-14-5	513
2857-97-8	504	3145-88-8	243	3564-26-9	533
2861-02-1	500	3147-14-6	295	3575-31-3	365
2861-99-6	215, 755	3150-24-1	830	3588-17-8	189
2869-83-2	453	3153-26-2	660	3588-91-8	316
2892-51-5	255	3167-49-5	405	3603-99-4	241
2916-68-9	545	3173-56-6	281	3604-87-3	849
2916-76-9	518	3182-93-2	793	3604-90-8	839
2922-54-5	652	3182-95-4	370	3608-75-1	465
2923-16-2	647	3184-13-2	793	3612-20-2	411
2923-18-4	654	3196-73-4	774	3615-37-0	822
2923-28-6	528, 709	3207-12-3	116	3615-41-6	831
2935-35-5	794	3211-76-5	795	3618-43-7	498
2937-50-0	113	3217-94-5	236	3625-57-8	466
2949-92-0	186	3219-63-4	545	3632-91-5	642
2950-43-8	583	3238-62-8	447	3634-56-8	515
2957-95-1	246	3240-34-4	306	3646-73-9	823
2961-50-4	407	3244-41-5	526	3680-02-2	189
2963-79-3	416	3251-23-8	573	3682-32-4	364
2966-50-9	649	3264-82-2	696	3682-35-7	494
2978-58-7	183	3277-26-7	535	3688-79-7	346
2978-59-8	183	3279-54-7	512	3699-66-9	540
2985-59-3	340	3282-30-2	197	3700-67-2	153, 753
2996-92-1	523	3282-56-2	294	3709-18-0	493
3001-72-7	427	3282-73-3	142, 753	3710-27-8	644
3002-78-6	460	3284-79-5	301	3713-25-5	826
3007-31-6	141	3287-99-8	279	3721-95-7	227
3008-40-0	239	3289-28-9	247	3722-51-8	449
3011-89-0	300	3307-39-9	301	3724-43-4	324
3012-37-1	283	3320-87-4	363	3733-85-5	468
3016-19-1	113	3326-32-7	53, 440	3734-67-6	523
3018-12-0	140	3350-30-9	233	3737-95-9	295
3019-22-5	251	3375-31-3	698	3739-30-8	172
3022-16-0	285	3376-24-7	294	3739-38-6	369
3024-83-7	783	3378-72-1	280	3739-64-8	113

3740-52-1	362	4156-16-5	439	4552-64-1	491
3767-28-0	830	4199-10-4	895	4556-23-4	479
3775-72-2	776	4199-88-6	466	4562-27-0	448
3790-78-1	189	4205-91-8	867	4568-71-2	411
3806-59-5	233	4206-67-1	515, 546	4592-94-3	288
3810-74-0	898	4221-98-1	254	4602-84-0	876
3811-04-9	599	4223-11-4	218	4611-62-5	438
3811-73-2	516	4224-70-8	119	4618-18-2	827
3829-86-5	470	4233-33-4	472	4627-10-5	750
3856-25-5	227	4252-56-6	496	4628-94-8	503
3857-36-1	491	4254-02-8	235	4630-80-2	236
3859-41-4	240	4254-14-2	197	4636-73-1	212
3866-79-3	333, 392	4254-15-3	197	4645-32-3	554
3871-20-3	548	4264-83-9	512	4648-54-8	543
3881-33-2	403	4265-07-0	892	4656-81-9	207, 754
3886-69-9	349	4271-30-1	775	4659-45-4	314
3886-70-2	357	4272-74-6	796	4672-38-2	527
3902-71-4	494	4282-31-4	491	4675-64-3	645
3930-19-6	898	4286-15-1	371	4693-91-8	347
3930-83-4	344	4294-16-0	862	4693-92-9	246
3952-78-1	401	4316-93-2	429	4695-62-9	249
3956-07-8	344	4328-13-6	207, 754	4696-76-8	882
3966-30-1	270	4329-75-3	436	4707-71-5	367
3966-32-3	347	4337-75-1	515	4712-70-3	414
3970-21-6	181	4350-09-8	790	4733-39-5	432
3978-81-2	417	4360-12-7	401, 498	4746-97-8	231
3982-91-0	598	4368-51-8	206	4781-83-3	450
3993-78-0	403	4370-62-1	868	4786-20-3	135
4001-73-4	287	4373-34-6	401	4792-83-0	271
4023-02-3	477	4375-83-1	729	4819-77-6	212
4023-65-8	110	4403-89-8	500	4837-88-1	352
4033-27-6	872	4403-90-1	500	4839-46-7	150
4036-30-1	375	4408-64-4	731, 732, 744	4841-91-2	237
4039-32-1	514	4410-48-4	401	4845-50-5	431
4041-95-6	140	4411-25-0	222	4845-99-2	865
4042-36-8	795	4411-26-1	222	4874-17-3	774
4043-87-2	473	4414-88-4	409	4874-65-1	248
4044-65-9	372	4419-47-0	655	4900-63-4	347
4045-44-7	253, 686	4426-47-5	505	4901-51-3	382
4050-45-7	170	4430-20-0	301	4916-57-8	436
4052-59-9	778	4431-00-9	271	4930-98-7	445
4054-38-0	228	4432-31-9	465	4942-47-6	220
4067-16-7	194	4436-27-5	377	4961-10-4	213
4071-88-9	513	4436-74-2	123	4984-22-9	200
4075-81-4	632	4441-63-8	229	4989-82-1	634, 648, 660
4088-22-6	153, 753	4466-59-5	315	4984-82-1	673, 721, 743
4097-89-6	216	4472-44-0	420	5006-66-6	448
4097-89-6	548	4477-28-5	395	5015-75-8	529
4105-21-9	471	4498-67-3	450	5051-62-7	337
4109-96-0	539	4511-42-6	432	5056-14-4	839
4111-54-0	641	4524-93-0	235	5057-98-7	238
4113-06-8	212	4525-33-1	149	5057-99-8	238
4113-12-6	131	4542-57-8	155	5058-42-4	640
4114-31-2	159	4547-57-3	293	5061-35-8	772
4128-17-0	876	4548-45-2	421	5070-13-3	286
4130-42-1	311	4551-15-9	522	5076-82-4	795

5080-50-2	772	5451-09-2	778	5850-86-2	521
5098-14-6	115	5451-40-1	429	5854-52-4	222
5103-96-8	896	5452-35-7	229	5856-62-2	115
5108-96-3	481	5455-59-4	359	5856-63-3	115
5122-82-7	221	5459-00-7	241	5892-11-5	238, 865
5130-24-5	218	5460-29-7	292	5897-76-7	267
5133-19-7	817	5464-70-0	333	5908-77-0	628
5137-55-3	752	5466-22-8	114	5908-94-1	270
5138-90-9	529	5466-84-2	363	5913-13-3	233
5138-92-1	529	5468-37-1	261	5922-92-9	207, 754
5138-93-2	529	5469-45-4	258	5930-72-3	510
5144-89-8	470	5470-11-1	583	5934-29-2	789
5153-67-3	364	5470-18-8	421	5936-98-1	538
5162-02-1	894	5470-96-2	482	5937-77-9	779
5174-67-4	496	5471-63-6	435	5949-05-3	135
5182-30-9	553	5490-27-7	873	5949-24-6	423
5188-07-8	531	5515-64-0	231	5949-29-1	135
5266-20-6	467	5522-66-7	895	5956-60-5	411
5267-34-5	330	5523-19-3	565	5959-52-4	266
5267-64-1	370	5538-51-2	261	5965-83-3	381
5283-66-9	520	5539-53-7	110	5967-42-0	346
5289-74-7	849	5541-93-5	892	5967-73-7	884
5292-43-3	125	5543-57-7	906	5968-90-1	497
5296-64-0	261	5543-58-8	906	5970-45-6	661
5310-10-1	252	5558-66-7	330	5977-14-0	105
5321-87-8	246	5593-70-4	655	5978-75-6	264
5324-84-5	531	5600-21-5	403	5979-01-1	906
5326-23-8	421	5614-37-9	241	5979-11-3	781
5327-44-6	325	5625-37-6	474	5995-86-8	336
5328-01-8	332	5635-50-7	285	5995-99-5	824
5329-14-6	617	5657-17-0	161	5996-00-9	824
5336-08-3	832	5662-95-3	236	6000-43-7	787
5341-61-7	580	5673-07-4	320	6001-51-0	285
5341-95-7	121	5676-48-2	383	6003-09-4	112
5343-41-9	531	5680-79-5	787	6003-10-7	112
5344-58-1	224	5704-04-1	797	6003-18-5	854
5348-42-5	307	5707-04-0	501	6004-24-6	419
5351-69-9	375	5735-89-3	897	6007-26-7	458
5391-88-8	290	5743-04-4	630	6009-70-7	627
5392-28-9	486	5743-26-0	631	6011-14-9	114
5392-67-6	505, 630	5743-36-2	631	6018-19-5	650
5394-63-8	493	5752-40-9	144, 150	6018-89-9	644
5396-89-4	279	5762-56-1	216	6020-87-7	782
5396-98-5	739	5763-44-0	237	6025-53-2	908
5398-44-7	430	5763-49-5	237	6027-13-0	789
5407-87-4	403	5765-84-0	845	6032-29-7	195
5412-98-6	205	5776-72-7	285	6034-57-7	399
5415-44-1	494	5785-66-0	306	6044-68-4	111
5419-55-6	541, 551	5794-04-7	225	6046-93-1	633
5427-03-2	311	5794-13-8	779	6050-13-1	328
5428-64-8	469	5796-17-8	784	6050-81-3	858
5433-88-7	275	5798-49-2	286	6055-52-3	170
5437-38-7	352	5798-94-7	292	6055-71-6	111
5437-45-6	280	5802-65-3	236	6057-60-9	285
5439-14-5	430	5807-14-7	492	6057-90-5	774
5440-00-6	426	5808-22-0	512	6066-82-6	172
5442-44-4	302	5838-82-4	211, 755	6068-62-8	166

6068-72-0	306	6303-58-8	369	6920-22-5	170
6074-84-6	654	6306-07-6	258	6921-35-3	433
6076-30-8	874	6315-52-2	333	6929-40-4	864
6078-64-4	785	6318-55-4	110	6939-89-5	381
6080-33-7	484	6325-93-5	359	6948-88-5	355
6080-56-4	640	6326-79-0	415	6950-84-1	357
6088-51-3	341	6358-69-6	553	6964-21-2	490
6090-09-1	219	6359-98-4	523	6967-12-0	404
6091-44-7	474	6363-53-7	828	6974-32-9	816
6091-45-8	474	6380-21-8	377	6976-37-0	118
6091-58-3	196	6381-59-5	646	6983-79-5	837
6094-40-2	474	6381-92-6	636	6990-27-8	729
6099-90-7	376	6399-81-1	549	6995-79-5	231
6100-05-6	645	6402-23-9	426, 814	6996-92-5	501
6100-20-5	646	6402-36-4	209	6998-60-3	897
6100-28-3	854	6409-77-4	520	7000-27-3	829
6106-07-6	796	6410-10-2	365	7048-02-4	789
6106-24-7	654	6418-56-0	508	7048-04-6	783
6108-23-2	641	6456-74-2	787	7057-27-4	872
6112-76-1	457	6459-94-5	519	7059-24-7	53, 866
6119-70-6	482	6476-37-5	736	7065-46-5	123
6132-04-3	660	6484-52-2	559	7080-50-4	506
6136-37-4	464	6487-48-5	646	7087-68-5	147, 414
6138-23-4	834	6505-50-6	642	7091-44-3	874
6138-41-6	492	6528-53-6	378	7114-03-6	350
6146-52-7	466	6532-76-9	790	7148-03-0	389
6147-53-1	633	6533-73-9	620	7188-38-7	127
6148-75-0	533	6554-98-9	342	7205-91-6	300
6151-25-3	481	6556-12-3	826	7205-98-3	300
6152-43-8	375	6575-24-2	315	7212-44-4	189
6153-39-5	365	6586-05-6	644	7216-56-0	113
6153-56-6	193	6600-40-4	793	7235-40-7	838
6153-64-6	898	6605-21-6	820	7250-67-1	420
6156-78-1	642	6609-64-9	506	7295-85-4	419
6159-66-6	902	6627-34-5	314	7296-64-2	823
6160-12-9	484	6627-72-1	224	7297-25-8	822
6160-65-2	489	6627-89-0	294	7322-88-5	259
6163-58-2	551	6627-91-4	296	7324-05-2	774
6170-22-5	854	6628-04-2	405	7331-34-4	318
6183-68-2	482	6628-37-1	531, 652	7343-51-3	172
6184-17-4	877	6635-20-7	364	7356-11-8	352
6192-52-5	387	6635-29-6	240, 243, 244	7364-25-2	450
6203-18-5	321	6651-43-0	546	7365-44-8	216
6211-16-1	793	6655-27-2	345	7365-45-9	446
6217-54-5	153	6674-22-2	427	7365-82-4	104
6221-55-2	223	6699-20-3	878	7369-44-0	303
6232-88-8	292, 295	6708-35-6	905	7379-35-3	422
6271-79-0	307	6722-15-2	458	7393-43-3	654
6278-91-7	282	6833-84-7	889	7397-62-8	126
6283-24-5	627	6833-86-7	627	7411-49-6	307
6283-74-5	136	6834-92-0	614	7413-60-7	787
6284-40-8	829	6866-70-2	538	7422-32-4	536
6293-20-5	781	6892-68-8	153	7439-88-5	585
6294-89-9	185	6893-02-3	797	7439-89-6	585
6297-03-6	191	6895-43-8	839	7439-91-0	585
6303-21-5	583	6915-15-7	178	7439-93-2	586

7439-95-4	588	7558-80-7	611	7705-08-0	575
7439-97-6	589	7560-83-0	246, 698	7707-57-5	424
7439-98-7	590	7562-61-0	902	7710-35-2	148
7440-01-9	592	7569-58-6	299	7718-54-9	592
7440-02-0	663	7569-60-0	299	7719-09-7	620
7440-05-3	664	7569-87-1	299	7719-12-2	598
7440-06-4	665	7580-67-8	587	7721-01-9	619
7440-15-5	667	7585-39-9	819	7722-43-7	563
7440-16-6	669	7597-56-0	162	7722-64-7	602
7440-18-8	670	7598-10-9	290	7722-76-1	558
7440-22-4	606	7601-89-0	615	7722-84-1	582
7440-23-5	609	7601-90-3	595	7722-86-3	603
7440-31-5	621	7606-79-3	773	7723-14-0	597
7440-33-7	622	7612-98-8	320	7724-15-4	388
7440-37-1	561	7631-86-9	606	7726-95-6	565
7440-39-3	562	7631-90-5	609	7727-15-3	556
7440-43-9	565	7631-94-9	611	7727-18-6	622
7440-44-0	131, 664, 924	7631-99-4	614	7727-21-1	602
7440-48-4	934	7632-00-0	614	7727-37-9	594
7440-55-3	577	7637-07-2	565	7727-54-0	559
7440-56-4	577	7641-77-2	249, 338	7727-73-3	615
7440-57-5	578	7643-75-6	818	7732-18-5	623
7440-59-7	579	7646-69-7	612	7733-02-0	625
7440-62-2	622	7646-78-8	617	7749-47-5	404
7440-66-6	623, 938	7646-79-9	572	7757-79-1	602
7440-69-9	564	7646-85-7	623	7757-82-6	615
7440-70-2	566	7646-93-7	599	7757-83-7	616
7440-74-6	584	7647-01-0	581, 582	7757-93-9	591
7446-07-3	619	7647-10-1	595, 743	7758-01-2	599
7446-08-4	605	7647-14-5	610	7758-02-3	599
7446-09-5	618	7647-15-6	610	7758-05-6	601
7446-18-6	620	7647-17-8	569	7758-09-0	602
7446-19-7	625	7647-73-6	133	7758-19-2	610
7446-20-0	625	7652-46-2	772	7758-29-4	616
7446-70-0	556	7658-08-4	820	7758-89-6	573
7447-39-4	573	7664-38-2	597	7758-94-3	576
7447-40-7	599	7664-39-3	581, 582	7758-95-4	585
7447-41-8	587, 743	7664-41-7	557	7758-98-7	573
7449-74-3	518	7664-93-9	618	7761-88-8	607
7460-82-4	143	7669-54-7	359	7772-98-7	616
7462-62-6	192	7677-21-8	208	7772-99-8	617
7464-93-9	494	7677-24-9	543	7773-01-5	589
7480-35-5	264	7681-11-0	601	7773-06-0	559
7486-35-3	661	7681-49-4	611	7774-29-0	589
7487-88-9	588	7681-55-2	613	7774-34-7	566
7487-94-7	589	7681-65-4	573	7775-09-9	610
7488-54-2	604	7681-82-5	613	7775-41-9	607
7488-99-5	838	7685-44-1	774	7778-50-9	600
7491-74-9	893	7688-21-3	170	7778-74-7	602
7493-95-0	829	7688-25-7	726	7778-77-0	600
7516-48-5	883	7695-91-2	905	7778-80-5	603
7529-22-8	459	7697-37-2	593	7781-49-9	452
7542-37-2	891	7699-35-6	496	7782-24-3	374
7550-35-8	586	7699-43-6	625	7782-26-5	374
7550-45-0	621	7699-45-8	623	7782-39-0	574
7553-56-2	584	7704-34-9	618	7782-40-3	932
7558-79-4	575	7705-07-9	621	7782-41-4	576

7782-42-5	579, 929, 931	7787-35-1	562	7791-18-6	588
7782-44-7	595	7787-38-4	563	7791-20-0	592
7782-49-2	605	7787-47-5	563	7791-25-2	619
7782-50-5	570	7787-50-0	564, 628	7791-28-8	562
7782-61-8	575	7787-55-5	564	7803-49-8	583
7782-63-0	576	7787-56-6	564	7803-51-2	523, 596
7782-68-5	584	7787-60-2	564	7803-55-6	560
7782-79-8	191, 580	7787-69-1	569	7803-57-8	580
7782-86-7	589	7787-70-4	573	7803-58-9	618
7782-89-0	586	7788-98-9	558	7803-63-6	557
7782-91-4	591	7788-99-0	571	7803-65-8	558
7782-92-5	609	7789-00-6	600	7803-68-1	619
7783-00-8	605	7789-02-8	571	8002-36-6	834
7783-06-4	583	7789-06-2	617	8002-43-5	882
7783-20-2	559	7789-09-5	558	8003-05-2	644
7783-21-3	559	7789-12-0	611	8004-87-3	353
7783-28-0	575	7789-17-5	570	8008-20-6	177
7783-34-8	589	7789-18-6	570	8012-95-1	193
7783-36-0	589	7789-20-0	574	8014-95-7	619
7783-49-5	624	7789-24-4	587	8015-61-0	817
7783-56-4	560	7789-29-9	601	8032-32-4	196
7783-66-6	584	7789-30-2	565	8048-52-0	53, 400
7783-70-2	560	7789-33-5	584	8052-16-2	858
7783-77-9	590	7789-38-0	610	8063-24-9	400
7783-81-5	622	7789-39-1	604	9000-01-5	934
7783-83-7	558	7789-40-4	620	9000-07-1	819
7783-85-9	558	7789-41-5	566	9000-69-5	830
7783-89-3	607	7789-42-6	565	9001-03-0	800
7783-90-6	607	7789-43-7	572	9001-12-1	802
7783-92-8	607	7789-45-9	572	9001-16-5	803
7783-93-9	608	7789-47-1	589	9001-26-7	812
7783-95-1	607	7789-48-2	588	9001-32-5	804
7783-97-3	607	7789-60-8	598	9001-37-0	805
7783-98-4	608	7789-68-6	621	9001-40-5	805
7783-99-5	608	7789-69-7	597	9001-50-7	806
7784-00-1	608, 694	7789-77-7	591	9001-54-1	807
7784-01-2	607	7789-80-2	567	9001-59-6	812
7784-17-0	556	7790-21-8	602	9001-60-9	808
7784-18-4	556	7790-28-5	614	9001-63-2	809
7784-24-9	556	7790-30-9	620	9001-73-4	810
7784-26-1	556	7790-32-1	588	9001-75-6	810
7784-27-2	556	7790-44-5	560	9001-78-9	810
7784-29-4	556	7790-47-8	617	9001-99-4	813
7784-31-8	556	7790-60-5	604	9002-03-3	803
7784-33-0	561	7790-69-4	587	9002-04-4	814
7784-34-1	561	7790-79-6	566	9002-68-0	805
7784-45-4	561	7790-80-9	566	9002-81-2	197
7785-20-8	559	7790-84-3	566	9002-88-4	197
7785-23-1	607	7790-91-2	570	9003-04-7	653 <sup>a</sup>
7785-26-4	255	7790-94-5	507, 570	9003-20-7	197
7785-70-8	255	7790-98-9	559	9003-21-8	197
7785-84-4	616	7790-99-0	584	9003-39-8	474
7787-20-4	249	7791-03-9	587	9003-53-6	377
7787-32-8	562	7791-11-9	604	9004-02-8	808
7787-33-9	562	7791-12-0	620	9004-07-3	801
7787-34-0	562	7791-13-1	572	9004-32-4	650

9004-54-0	821	10025-67-9	619	10101-63-0	585
9005-32-7	817	10025-73-7	570	10101-89-0	621
9005-38-3	650	10025-77-1	575, 712	10101-97-0	593
9005-49-6	826	10025-78-1	725	10101-98-1	593
9005-79-2	826	10025-78-2	538, 676	10102-17-7	616
9005-82-7	817	10025-82-8	584	10102-18-8	615
9005-84-9	833	10025-87-3	597	10102-40-6	614
9007-31-2	811	10025-91-9	560	10102-43-9	594
9007-43-6	802	10025-98-6	603	10102-45-1	620
9007-73-2	804	10025-99-7	603	10102-68-8	567
9008-02-0	806	10026-04-7	606	10108-64-2	566
9012-31-1	798	10026-06-9	617	10118-76-0	567
9012-33-3	798	10026-08-1	621	10124-36-4	566
9013-18-7	798	10026-11-6	625	10124-41-1	567
9013-38-1	804	10026-12-7	593	10124-43-3	572
9014-01-1	814	10026-13-8	597	10124-65-9	646
9014-56-6	806	10026-22-9	572	10138-04-2	558
9015-82-1	800	10026-24-1	572	10139-47-6	624
9015-83-2	811	10028-17-8	622	10170-69-1	588
9025-26-7	801	10028-22-5	575	10181-37-0	472
9025-52-9	815	10030-85-0	831	10191-18-1	118
9026-30-6	799	10031-20-6	588	10191-41-0	900
9027-03-6	815	10031-22-8	585	10199-89-0	421
9027-60-5	798	10031-26-2	575	10210-64-7	628
9027-73-0	810	10031-43-3	573	10210-68-1	686, 934
9028-00-6	801	10033-92-8	149, 432	10213-10-2	617
9028-38-0	807	10034-76-1	567	10217-52-4	580
9028-49-3	798	10034-81-8	588	10230-26-9	245
9029-22-6	815	10034-82-9	610	10236-47-2	829
9030-66-4	806	10034-85-2	581, 582	10248-30-3	455
9030-74-4	807	10034-93-2	580	10281-56-8	135
9031-11-2	805	10034-96-5	589	10294-26-5	608
9031-37-2	801	10035-04-8	566	10294-28-7	578
9031-94-1	803	10035-10-6	581	10294-29-8	578, 732
9031-96-3	812	10038-98-9	578	10294-31-2	579
9034-42-8	792	10039-54-0	583	10294-34-5	565
9038-94-2	810	10039-56-2	613	10294-38-9	562
9039-53-6	815	10041-06-2	450	10294-46-9	573
9041-08-1	826	10042-76-9	617	10294-54-9	570
9046-40-6	830	10043-01-3	556	10294-66-3	604
9047-17-0	810	10043-35-3	564	10303-95-4	242
9047-22-7	801	10043-52-4	566	10315-85-2	170
9048-46-8	799	10045-89-3	558	10318-18-0	783
9050-98-0	261	10048-13-2	898	10323-20-3	818
9054-89-1	802	10048-95-0	609	10326-27-9	562
9062-61-7	788	10049-05-5	571	10334-26-6	225
9064-57-7	819	10049-07-7	604	10343-99-4	425
9066-59-5	809	10049-07-7	705	10361-79-2	604
9068-38-6	813	10049-08-8	706	10373-78-1	225
9074-11-7	803	10049-21-5	611	10377-48-7	587
9075-65-4	806	10075-50-0	415	10377-51-2	587, 744
10009-70-8	120	10075-62-4	319	10377-58-9	588
10016-20-3	819	10075-63-5	319	10381-37-0	621
10022-31-8	563	10097-28-6	606	10381-75-6	416
10022-68-1	566	10097-84-4	486	10388-19-9	344
10024-97-2	594	10099-74-8	586	10389-65-8	780
10025-65-7	599	10101-41-4	567	10397-30-5	352

10415-75-5	589	12060-05-8	667	13204-98-3	790
10416-59-8	503	12078-25-0	636	13207-77-3	876
10420-89-0	357	12080-32-9	685	13242-55-2	818
10424-65-4	207	12081-12-8	590, 687	13250-12-9	124
10441-45-9	340	12082-08-5	628	13257-51-7	643
10450-59-6	568	12092-47-6	677, 703, 747	13257-81-3	546
10466-65-6	603	12093-10-6	638	13268-42-3	627
10466-65-6	721	12099-10-4	206	13269-52-8	170
10467-46-6	840	12101-02-9	757, 760	13274-43-6	463
10476-81-0	617	12107-56-1	684	13280-60-9	266
10517-44-9	137	12112-67-3	414, 689, 691	13283-01-7	622
10519-96-7	526	12122-67-7	554	13291-61-7	230
10523-68-9	221	12124-97-9	557	13292-46-1	897
10527-16-9	297	12125-02-9	557	13323-81-4	372
10531-50-7	375	12130-66-4	701	13327-32-7	563
10534-87-9	572	12137-21-2	666	13346-19-0	588
10534-89-1	579	12143-03-2	667	13361-64-3	546
10544-72-6	575	12145-47-0	684	13362-78-2	436
10545-99-0	618	12145-48-1	685	13388-75-5	320
10550-58-0	131	12148-71-9	691	13389-42-9	192
10569-72-9	777	12150-46-8	630	13396-41-3	616
10593-83-6	838	12170-97-7	685	13400-13-0	570
10595-51-4	352	12192-57-3	627	13408-09-8	512
10597-89-4	886	12230-71-6	562	13410-01-0	615
10601-80-6	160	12245-39-5	703	13411-42-2	544
11016-17-4	53	12246-51-4	690	13419-61-9	529
11018-89-6	853	12257-42-0	629, 704, 747	13429-83-9	224
11018-93-2	833	12284-43-4	800	13435-12-6	542
11024-24-1	821	12333-90-4	230	13446-03-2	588
11025-94-8	812	12354-84-6	686	13446-09-8	558
11029-02-0	875	12354-85-7	686	13446-18-9	588
11029-61-1	788	12389-34-3	542	13446-34-9	589
11032-88-5	802	12427-38-2	516	13446-67-8	604
11075-17-5	800	12501-23-4	598	13446-71-4	604
11089-65-9	901	12585-29-4	809	13450-90-3	577
11096-37-0	815	12671-19-1	809	13450-92-5	578
11110-52-4	613, 745	12676-20-9	835	13453-30-0	620, 235
11113-62-5	875	12682-85-8	934	13453-38-8	235
11120-25-5	560	13007-92-6	571	13453-40-2	620
11121-48-5	528	13011-54-6	609	13454-84-7	570
11135-81-2	615	13057-50-6	831	13454-94-9	568
11171-61-9	135	13058-04-3	876	13455-01-1	598
12012-95-2	626, 671, 721	13059-60-4	787	13455-20-4	600
12026-06-1	620	13071-11-9	895	13455-24-8	599
12027-06-4	559	13072-69-0	223	13462-88-9	592
12027-43-9	606	13073-35-3	785	13462-93-6	558
12035-72-2	593	13100-82-8	783	13463-40-6	585, 634
12036-09-8	668	13106-76-8	559	13464-82-9	584
12036-10-1	605, 671	13113-71-8	270	13464-92-1	565 <sup>†</sup>
12037-01-3	619	13114-62-0	357	13465-95-7	563
12037-29-5	604	13126-12-0	604	13472-35-0	611
12038-63-0	669	13138-48-2	323	13472-36-1	615
12038-67-4	669	13149-00-3	230	13473-90-0	556
12044-50-7	561	13170-43-9	543, 546	13476-99-8	660
12053-18-8	663	13183-20-5	375	13477-00-4	562
12054-85-2	559, 695	13194-68-8	265	13477-34-4	567

13477-89-9	591	13821-08-4	566	14284-92-5	702
13478-00-7	592	13825-74-6	621	14284-93-6	647, 705, 707
13478-10-9	576	13826-27-2	723	14293-70-0	602
13478-18-7	590	13826-83-0	560	14306-25-3	653
13478-33-6	572	13842-46-1	602	14320-04-8	661
13494-80-9	619	13845-17-5	562	14324-55-1	554
13494-91-2	577	13862-28-7	525	14338-32-0	421
13510-42-4	604	13874-02-7	678, 679	14350-52-8	216
13520-61-1	592	13889-98-0	398	14358-33-9	780
13520-83-7	622	13894-21-8	334	14370-45-7	253
13520-92-8	625	13929-35-6	897	14376-79-5	256
13537-24-1	575	13933-23-8	576	14389-77-6	778
13545-04-5	152	13939-06-5	590	14401-91-3	473
13548-38-4	571	13963-57-0	626	14402-88-1	637
13548-43-1	570	13963-58-1	600	14444-77-0	316
13550-47-5	568	14002-51-8	284	14459-29-1	880
13566-22-8	559	14019-81-9	541	14459-95-1	600
13567-05-7	818	14024-17-0	639, 689, 692	14463-68-4	520, 887
13568-33-7	587	14024-18-1	689, 692	14481-29-9	558
13573-16-5	559	14024-41-1	601	14508-49-7	421
13587-99-4	563, 564	14024-48-7	633	14532-22-9	604
13590-82-4	569	14024-58-9	642	14540-59-1	549
13596-22-0	572	14024-61-4	698	14592-56-4	672
13596-46-8	572	14024-63-6	661	14618-80-5	411
13598-36-2	596	14040-11-0	622	14620-72-5	148
13598-65-7	721	14055-02-8	644	14630-40-1	503
13600-98-1	612	14073-97-3	251	14631-20-0	857
13601-11-1	601	14074-80-7	661	14637-34-4	499
13601-19-9	611	14075-53-7	603	14649-03-7	373
13601-63-3	566	14090-77-8	275	14691-89-5	397
13623-25-1	347	14098-24-9	409	14694-95-2	679
13637-61-1	624	14098-44-3	409	14698-29-4	468
13637-68-8	590, 687	14103-61-8	328	14704-14-4	517
13716-10-4	508	14104-20-2	648, 707	14717-29-4	737
13716-10-4	722, 734, 735	14121-36-9	486	14724-63-1	883
13716-12-6	538	14121-89-2	417	14737-80-5	382
13716-45-5	509	14165-55-0	639	14777-27-6	182
13718-50-8	562	14166-21-3	230	14786-51-7	900
13719-97-6	293	14167-18-1	630	14871-79-5	562
13737-04-7	546	14172-90-8	633	14898-67-0	604, 706
13746-66-2	600	14172-92-0	644	14906-55-9	440
13755-29-8	611	14174-09-5	428	14910-06-6	480
13756-40-6	345	14187-32-7	427	14930-96-2	871
13762-51-1	599	14217-21-1	611	14937-45-2	754
13765-03-2	587	14220-64-5	672	14977-61-8	571
13768-11-1	595, 721	14221-01-3	708	14984-39-5	824
13770-18-8	573	14242-05-8	608	14990-02-4	437
13776-84-6	616	14243-64-2	679, 708	14999-66-7	210
13779-37-2	547	14243-64-2	709, 710	15014-25-2	309
13780-06-8	567	14260-41-4	295	15016-43-0	553
13780-42-2	577	14264-16-5	696	15023-99-1	310
13795-24-9	282	14268-66-7	350	15042-77-0	294
13811-71-7	146	14271-10-9	303	15123-90-7	533
13812-88-9	566	14277-97-5	875	15170-57-7	645
13813-24-6	592	14282-91-8	580	15206-55-0	349
13815-94-6	685, 706	14283-07-9	587	15207-30-4	788
13821-06-2	562	14284-36-7	556	15231-91-1	291

15244-36-7	589	16495-13-9	411	17261-28-8	733, 741
15284-15-8	884	16523-54-9	505, 506, 737	17279-31-1	323
15318-45-3	900	16526-68-4	231	17325-85-8	282
15321-51-4	585	16564-43-5	652	17331-79-2	780
15356-60-2	760	16574-43-9	292	17334-55-3	224
15421-51-9	882	16586-35-7	134	17338-96-4	464
15435-71-9	649, 681	16626-02-1	786	17354-14-2	381
15448-47-2	475	16635-95-3	739	17372-87-1	637
15489-90-4	880	16674-78-5	641	17397-89-6	130
15500-66-0	853	16676-29-2	887	17406-45-0	901
15525-13-0	208	16691-43-3	404	17422-32-1	420
15529-49-4	699, 707	16709-86-7	506	17422-74-1	877
15537-71-0	858	16713-15-8	263	17430-98-7	233
15548-52-4	901	16713-66-9	236	17435-72-2	158
15596-07-3	550	16733-97-4	634	17442-18-1	655
15603-10-8	640	16751-02-3	879	17448-65-6	154
15610-76-1	572	16774-21-3	693	17455-13-9	423
15629-92-2	198	16774-21-3	568, 569	17463-35-3	193
15630-89-4	614, 695	16812-54-7	593	17472-78-5	843
15635-87-7	742	16834-13-2	488	17501-44-9	661
15663-27-1	575	16836-95-6	528	17524-05-9	695
15667-21-7	438	16849-91-5	368	17534-15-5	274
15699-18-0	559	16853-85-3	586	17539-31-0	385
15743-47-5	632	16856-13-6	780	17549-30-3	515
15780-02-6	209	16871-60-6	601	17564-64-6	300
15825-70-4	828	16871-90-2	600	17575-20-1	827
15861-24-2	424	16872-11-0	576, 749	17575-21-2	827
15892-23-6	124	16883-45-7	534	17575-22-3	827
15910-23-3	847	16893-85-9	612	17579-79-2	351
15914-84-8	356	16903-35-8	579, 679, 732	17579-99-6	518
15936-16-0	465	16919-19-0	558	17615-97-3	154
15956-28-2	702, 705	16919-58-7	579, 598, 666	17629-30-0	831
15962-47-7	772	16920-56-2	601	17640-15-2	183
15988-11-1	472	16921-30-5	601	17680-01-2	527
16009-13-5	880	16923-95-8	601	17685-52-8	621
16029-98-4	515	16940-66-2	609	17687-47-7	260
16039-53-5	661	16940-81-1	748	17697-12-0	501
16039-55-7	631	16940-92-4	558	17702-41-9	508
16052-42-9	252	16941-11-0	558	17702-41-9	574
16055-80-4	455	16941-12-1	579, 665	17728-88-0	513
16069-36-6	430	16941-92-7	585	17768-41-1	221
16088-62-3	475	16949-15-8	586	17833-43-1	829
16093-66-6	393	16961-25-4	678	17872-92-3	546
16156-59-5	374	16965-08-5	294	17873-08-4	523
16182-04-0	157	17016-83-0	759	17896-21-8	773
16202-15-6	395	17022-31-0	200	17927-65-0	556
16226-54-3	435	17026-42-5	138	18016-24-5	631
16266-99-2	836	17069-79-3	584	18027-10-6	533
16298-03-6	458	17084-13-8	601	18043-71-5	534
16355-00-3	372	17125-60-9	632	18069-17-5	184
16357-59-8	438	17140-60-2	632	18107-18-1	138, 543
16373-93-6	871	17146-86-0	780	18123-20-1	445
16409-44-2	878	17199-29-0	346	18125-46-7	786
16420-13-6	153	17224-73-6	238	18143-30-1	544
16454-60-7	592	17252-51-6	437	18162-48-6	505
16456-81-8	640	17257-71-5	348	18166-02-4	545

18194-24-6	874	19836-78-3	460, 468	21816-08-0	224
18197-26-7	651	19883-78-4	786	21820-51-9	893
18223-42-2	574	19891-74-8	840	21860-03-7	311
18243-41-9	546	19966-81-5	737	21908-53-2	589
18263-08-6	643	19980-43-9	546	21917-20-4	238
18263-25-7	119	20039-37-6	647	22037-28-1	415
18282-10-5	617	20049-65-4	576	22083-74-5	888
18293-53-3	542	20082-71-7	521	22115-41-9	290
18293-54-4	547	20108-77-4	774	22135-49-5	371
18293-71-5	543	20149-10-4	639	22147-81-5	435
18301-46-7	547	20230-89-1	215	22147-82-6	435
18378-89-7	885	20312-36-1	342	22147-83-7	435
18390-55-1	641	20350-15-6	864	22147-84-8	435
18395-90-9	508	20350-44-1	900	22150-76-1	863
18404-47-2	645	20357-25-9	319	22206-57-1	720
18422-05-4	859	20398-06-5	657, 691	22240-02-4	680
18467-77-1	822	20411-57-8	385	22282-75-3	440
18480-07-4	617	20439-47-8	243, 734	22282-76-4	440
18531-94-7	283, 712, 713	20440-95-3	332	22348-32-9	436, 717
18531-99-2	283, 712	20445-31-2	348	22373-78-0	653
18534-18-4	474	20445-33-4	348	22466-49-5	682
18546-37-7	820	20461-99-8	439	22483-09-6	114
18582-10-5	938	20462-00-4	438	22509-74-6	295
18636-55-0	323	20511-12-0	404	22608-11-3	870
18666-68-7	551	20562-02-1	832	22608-12-4	870
18685-01-3	370	20624-25-3	529	22685-96-8	911
18699-02-0	259	20667-12-3	608	22719-01-3	630
18709-01-8	232	20717-86-6	633	22809-37-6	120, 939
18714-16-4	471	20734-56-9	308	22810-82-2	490
18740-59-5	537	20734-57-0	308	22818-40-2	790
18742-02-4	414	20734-58-1	384	22990-77-8	403
18812-51-6	527	20756-92-7	553	23007-85-4	460
18868-43-4	590	20765-98-4	705	23110-15-8	441
18875-72-4	901	20816-12-0	595	23190-16-1	264
18883-66-4	833	20826-04-4	416	23220-52-2	784
18917-37-1	642	20830-75-5	822	23231-91-6	204
18997-57-4	829	20861-69-2	559	23235-01-0	795
19005-93-7	451	20902-45-8	891	23313-68-0	902
19010-79-8	631	20956-07-4	235	23398-40-5	530
19121-58-5	833	21106-96-7	861	23491-52-3	53
19132-06-0	122	21124-40-3	778	23519-77-9	661
19172-47-5	515, 938	21132-76-3	117	23537-25-9	783
19237-84-4	894	21151-51-9	121	23586-53-0	657
19341-57-2	859	21193-75-9	823	23897-15-6	541
19342-01-9	323	21205-91-4	503	23978-55-4	488
19444-84-9	171	21272-85-5	361	24009-17-4	839
19597-69-4	586	21286-54-4	226	24074-26-8	413, 414
19617-43-7	157	21291-99-6	209	24161-30-6	284
19694-02-1	378	21293-29-8	219	24221-86-1	875
19701-83-8	773	21302-43-2	403	24280-93-1	886
19718-36-6	602	21324-39-0	612	24295-03-2	399
19719-28-9	315	21436-03-3	243, 734	24346-00-7	860
19767-45-4	531	21679-31-2	680	24347-58-8	122
19780-67-7	163	21679-46-9	633	24424-99-5	139
19780-84-8	171	21715-90-2	250	24425-40-9	264
19810-31-2	282	21736-83-4	897	24448-89-3	447
19814-75-6	433, 751	21799-87-1	514	24463-19-2	299

24480-38-4	839	26444-49-5	512	29159-23-7	262
24544-04-5	331, 687	26472-41-3	250	29270-56-2	440
24589-37-3	338	26550-55-0	200	29420-49-3	525, 646
24589-78-4	518	26562-81-2	220	29427-58-5	185
24677-78-9	317	26585-94-4	840	29428-34-0	772
24686-78-0	474	26593-50-0	351	29548-30-9	876
24752-66-7	280	26614-21-1	285	29576-14-5	135
24815-24-5	483	26614-22-2	285	29617-66-1	134
24828-59-9	226	26628-22-8	609	29701-07-3	882
24830-94-2	774	26632-17-7	173	29826-91-3	733
24850-33-7	626	26651-23-0	378	29836-26-8	830
24857-20-3	509	26791-73-1	256	29841-69-8	677, 739
24887-75-0	842	26851-19-9	453	29870-99-3	654
24994-60-3	406	26908-82-7	261	29883-15-6	817
25014-41-9	197	27025-41-8	786	29892-37-3	678, 732
25015-63-8	747	27061-78-5	774	29943-42-8	487
25025-59-6	858	27072-45-3	440	29958-06-0	404
25037-45-0	915	27126-76-7	342	30108-39-5	541
25043-00-9	408	27215-10-7	502	30145-40-5	525
25067-59-8	474	27428-57-5	289	30334-71-5	874
25103-23-5	502, 552	27440-00-2	773	30345-24-5	464
25134-21-8	253	27496-82-2	350	30409-21-3	464
25144-64-3	368	27532-96-3	787	30471-18-2	698
25155-30-0	530	27565-41-9	153	30525-89-4	165
25232-41-1	474	27599-63-9	440	30587-85-0	533
25246-27-9	223	27738-96-1	132	30802-37-0	772
25249-54-1	474	27757-85-3	488	30983-70-1	333
25269-18-5	846	27779-29-9	251	31011-57-1	621
25278-25-0	683	27821-45-0	902	31111-21-4	603
25312-65-6	845	27854-88-2	480	31126-95-1	705
25316-59-0	283	27926-72-3	497	31250-06-3	488
25322-68-3	893	27988-67-6	901	31317-17-6	432
25379-20-8	471	28059-64-5	264	31362-50-2	780
25395-32-8	816	28110-26-1	273	31366-25-3	488
25465-65-0	251	28166-06-5	335	31469-15-5	517
25535-16-4	475	28169-46-2	327	31570-04-4	551
25542-62-5	158	28224-73-9	821	31597-37-0	243
25675-29-0	290	28240-09-7	865	31642-12-3	632
25704-18-1	523	28300-74-5	645	31857-42-8	427
25717-83-3	417	28393-20-6	843	31886-57-4	738
25779-13-9	372	28407-51-4	580	31886-58-5	738
25836-74-2	118	28469-92-3	315	31904-34-4	738
25895-60-7	611	28519-50-8	333, 387, 392	31915-82-9	175
25952-53-8	148	28529-34-2	773	32005-36-0	673
25955-51-6	691, 704	28553-12-0	328	32014-70-3	426
25997-59-5	819	28562-53-0	397	32161-06-1	398
26062-79-3	197	28643-94-9	779	32181-59-2	816
26078-96-6	654	28697-53-2	818	32248-43-4	133
26093-31-2	404	28728-55-4	149	32248-43-4	605, 707
26127-08-2	758	28736-68-7	639	32305-98-9	692
26164-26-1	347	28822-58-4	452	32315-10-9	119
26239-55-4	104, 474	28832-64-6	405	32354-50-0	704
26299-14-9	647	28903-20-0	516	32362-88-2	480
26305-75-9	733, 743	28920-43-6	335	32449-92-6	826
26386-88-9	511, 544	28994-41-4	282	32483-15-1	773
26412-87-3	533	29106-32-9	226	32503-27-8	203

32507-32-7	284	35231-44-8	53, 416	39382-08-6	892
32551-87-4	385	35293-35-7	300, 687	39416-48-3	479
32580-92-0	193	35455-24-4	773	39637-74-6	224
32638-88-3	527	35486-42-1	403	39637-99-5	348
32679-19-9	698	35556-70-8	892	39648-67-4	713
32685-69-1	790	35963-20-3	225	39665-12-8	791
32864-38-3	126	36042-94-1	506	39699-90-6	738
32899-41-5	891	36193-63-6	713	39717-48-1	464
33008-18-3	352	36362-97-7	274	39777-58-7	674, 725
33070-12-1	880	36394-75-9	108	39864-68-1	509
33100-27-5	423	36411-99-1	505	39961-95-0	243
33325-40-5	110	36520-39-5	408	40018-26-6	431
33375-06-3	373	36548-60-4	757	40064-34-4	474
33450-09-8	412	36548-61-5	757	40290-32-2	874
33628-84-1	780	36620-11-8	704	40499-83-0	449
33677-56-4	701	36635-61-7	389	40535-45-3	243
33725-74-5	533, 654	36653-82-4	131	40585-51-1	653
33745-47-0	295	37002-48-5	692	40615-36-9	320
33755-53-2	864	37026-85-0	210	40626-28-7	889
33850-87-2	210	37189-34-7	800	40698-91-7	701, 733
33884-43-4	414	37222-66-5	602	40716-66-3	189
33973-59-0	487	37278-10-7	811	40835-96-9	405
33998-38-8	306	37278-89-0	815	40835-97-0	503
34006-16-1	490	37299-04-0	660	40876-98-0	651
34017-20-4	773	37326-33-3	807	41024-90-2	283
34042-00-7	790	37342-97-5	711	41330-23-8	243
34122-28-6	265, 675, 747	37349-36-3	604	41345-70-4	110
34202-69-2	169	37350-66-6	624	41372-08-1	785
34255-08-8	840	37366-09-9	677, 707, 748	41372-20-7	862
34271-24-4	772	37577-07-4	889	41498-71-9	253
34275-72-4	570	37577-28-9	889	41661-47-6	414, 474
34301-54-7	938	37595-74-7	371	41667-95-2	430
34305-47-0	356	37718-11-9	477	41845-15-2	887
34316-15-9	419	37883-00-4	903	41864-22-6	419
34430-73-4	603	37933-66-7	833	41979-39-9	474
34464-40-9	175	37933-88-3	774	42151-56-4	350
34521-09-0	650	38099-82-0	878	42177-25-3	356
34622-08-7	694	38117-54-3	634, 757	42196-31-6	699
34713-94-5	182	38183-12-9	53, 441	42344-05-8	323
34822-90-7	635, 648	38222-83-2	428	42506-03-6	421
34830-11-0	660	38345-66-3	321	42740-82-9	747
34834-11-0	758	38353-09-2	448	42766-10-9	458
34836-53-8	508	38496-18-3	429	42852-95-9	829
34837-55-3	501	38533-41-4	858	42856-71-3	719, 720
34946-82-2	507	38630-75-0	251	42882-31-5	357
34994-50-8	896	38664-03-8	118	43189-45-3	778
35005-61-9	266	38894-81-4	838	44901-24-8	905
35011-47-3	449	38916-34-6	796	45497-73-2	268
35013-72-0	864	38928-89-1	644	46022-05-3	230
35015-46-4	704	39070-63-8	307	46292-93-7	375
35059-50-8	125	39145-52-3	773	47797-98-8	385
35061-50-8	821	39199-93-4	370	49637-59-4	509
35112-53-9	563	39208-14-5	636	49716-63-4	660
35132-20-8	739	39254-48-3	501	49759-20-8	400
35138-22-8	689, 692, 703	39262-22-1	226	50418-68-3	444
35138-23-9	689	39269-10-8	220	50471-44-8	496
35193-64-7	713	39319-11-4	549	50493-27-9	915

50525-27-4	660	54125-02-9	517	60459-08-7	572
50541-93-0	403	54162-90-2	223	60469-70-7	124, 293
50606-58-1	413	54249-88-6	803	60576-59-1	704
50696-65-6	637	54328-33-5	803	60633-21-8	735
50795-24-9	652	54656-96-1	480	60669-69-4	508
50812-37-8	806	54678-23-8	508	60676-86-0	606
50840-23-8	405	54816-44-5	630	60705-62-6	416
50887-69-9	467	55310-46-8	529	60835-69-0	466
50926-11-9	935	55441-72-0	494	60871-83-2	516
50936-05-5	492	55441-82-2	494	61183-76-4	652
50982-12-2	648, 684, 748	55453-09-3	518	61393-36-0	591
51019-43-3	259	55566-30-8	534	61475-31-8	249
51063-97-9	153	55599-62-7	808	61537-20-0	378
51076-95-0	134	55605-27-1	287	61550-02-5	546
51146-56-6	342	55870-50-3	224, 504	61717-82-6	445, 494
51146-57-7	342	55962-05-5	683, 917	61792-26-5	525
51206-73-6	420	56090-01-8	656, 658, 694	61886-02-0	703
51232-88-3	692	56320-90-2	570	61886-03-1	703
51364-51-3	709	56411-57-5	808	62087-82-5	222
51424-66-9	842	56484-13-0	661	62337-00-2	500
51481-60-8	886	56602-33-6	502, 721, 722	62418-65-9	540
51524-84-6	348	56660-19-6	630	62476-56-6	403
51569-18-7	519, 643	56680-61-5	316	62571-86-2	418
51877-53-3	642	56962-08-4	315	62613-82-5	890
51901-85-0	504	57044-25-4	443	62625-24-5	370
51905-34-1	641	57233-86-0	362	62648-72-0	566
51934-41-9	344	57266-56-5	223	62667-64-5	638
52093-29-5	693	57295-30-4	340	62697-73-8	792
52178-50-4	295	57605-95-5	356	62708-56-9	138
52225-20-4	902	57707-91-2	374	62758-12-7	513
52273-54-8	392	57794-08-8	230	63082-45-1	335
52351-64-1	660	57981-02-9	366	63231-67-4	606
52417-22-8	402	58123-62-9	719	63277-17-8	651
52429-99-9	551	58313-23-8	344	63363-93-9	262
52509-14-5	510	58477-85-3	137	63393-96-4	752
52522-40-4	709	58479-61-1	505	63395-36-8	629, 634
52665-69-7	632	58584-83-1	429	63401-04-7	435
52698-84-7	510	58616-58-3	691, 704	63401-12-7	447
52746-49-3	511	58632-95-4	293	63407-54-5	831
52766-61-2	223	58648-09-2	282	63409-48-3	891
52785-06-5	666	58801-34-6	632	63451-28-5	651
52809-07-1	795	58821-96-8	636	63451-29-6	527
52813-63-5	172	58823-95-3	879	63451-34-3	629
53022-70-1	694	58888-58-7	478	63451-35-4	514
53229-92-8	248	59025-03-5	108	63462-65-7	577
53317-87-6	710	59532-46-6	846	64241-57-2	794
53402-10-1	256	59557-02-7	240	64242-26-8	328
53408-96-1	466	59572-10-0	537	64315-16-8	779, 800
53518-15-3	407	59752-57-7	873	64443-05-6	683, 748
53563-56-7	285	59855-11-7	262	64550-88-5	832
53585-93-6	249	60046-49-3	256	64681-08-9	879
53801-63-1	624	60295-51-4	368	64732-91-8	878
53956-04-0	826	60310-07-8	352	64896-27-1	725
54010-71-8	879	60314-43-4	559	64896-28-2	122, 703, 725
54010-75-2	554	60434-71-1	198, 726	64918-47-4	774
54078-29-4	427	60456-23-7	443	64936-53-4	774

65012-73-9	703	71626-99-8	566	77287-29-7	183
65012-74-0	703	71807-57-3	794	77495-66-0	444
65039-08-9	756	72040-63-2	864	77882-24-7	510
65039-09-0	756	72120-52-6	901	78092-53-2	417
65058-23-3	778	72432-10-1	861	78338-22-4	490
65253-04-5	374	72497-47-3	809	78348-24-0	451, 719
65373-82-2	490	72497-48-4	809	78498-55-2	192
65604-89-9	216, 547, 548	72749-59-8	614, 695, 752	78508-96-0	447
66009-10-7	828	72845-94-4	53	78617-12-6	827
66019-20-3	500	72909-34-3	884	78905-09-6	546
66402-68-4	934	72914-19-3	734	78923-65-6	331
66455-27-4	810	73018-10-7	479	79004-97-0	531
66575-29-9	441	73183-34-3	729	79410-20-1	256
66640-86-6	864	73476-18-3	545	79790-89-9	704
66778-08-3	512	73487-00-0	630	79815-20-6	717, 719
67333-70-4	243	73500-82-0	418	79858-41-6	474
67373-56-2	537	73536-95-5	859	79917-90-1	756
67383-62-4	846	73573-88-3	868	80222-96-4	774
67392-80-7	846	73597-16-7	342	80403-59-4	652
67488-50-0	457	73871-24-6	692	80413-46-3	737
67881-59-8	704	74050-98-9	454	80695-16-5	629
67884-32-6	726	74090-15-6	236	80731-10-8	892
67884-58-7	704	74267-27-9	632	81290-20-2	547
67990-19-6	794	74677-08-0	509	81360-42-1	487
67999-30-8	224	74839-83-1	726	81484-86-8	537
68078-09-1	720	74839-84-2	725	81790-10-5	504
68168-23-0	819	74866-28-7	310	81927-55-1	283
68206-45-1	466	74892-62-9	704	82113-18-5	731
68252-28-8	890	75507-26-5	447	82113-65-3	676, 708, 730
68567-97-5	890	75513-72-3	642	82203-66-5	563
68696-18-4	205	75621-03-3	844	82410-32-0	878
68792-52-9	885	75640-87-8	712	82473-24-3	844
68831-78-7	835	75650-38-3	295	82495-75-8	545
68957-94-8	527	75708-92-8	877	82499-02-3	420
68971-82-4	417	75844-69-8	413	82737-61-9	502
69026-14-8	282	75853-08-6	367	82911-69-1	335
69056-38-8	899	75863-37-5	705	82923-76-0	719
69098-15-3	592	76058-33-8	334	83646-41-7	454
69150-58-9	401	76093-98-6	623	84210-26-4	880, 881
69221-14-3	712, 752, 867	76155-78-7	290	84227-47-4	410
69257-04-1	712, 752, 867	76186-04-4	759	84366-81-4	877
69340-58-5	469	76189-55-4	675, 714, 725	84665-66-7	643
69477-27-6	423	76189-56-5	684, 714, 725	85017-77-2	572
69815-49-2	889	76195-82-9	715	85051-91-8	528
69984-73-2	830	76231-37-3	397	85100-77-2	756
70024-90-7	799	76284-65-6	310	85196-70-9	133
70179-60-3	236	76422-63-4	503	85196-71-0	133
70197-13-6	656, 694	76496-51-0	270	85261-19-4	830
70268-36-9	467	76513-69-4	545	85261-20-7	819
70323-44-3	810	76571-64-7	772	85272-31-7	509
70445-62-4	429	76578-14-8	896	85362-85-2	285
70658-61-6	503	76584-70-8	200	85362-86-3	285
70693-62-8	445, 602, 747	76650-89-7	537	85618-20-8	827
70937-97-2	800	76836-02-7	532	85908-96-9	413
71264-71-6	703	76855-69-1	499	86046-11-9	719
71264-72-7	704	76899-34-8	629	86088-83-7	805
71626-98-7	567	77060-21-0	529	86632-33-9	725

86688-08-6	310	101055-56-5	410	119189-24-1	809
86703-52-8	238, 239	101055-57-6	410	120666-13-9	547
87243-73-0	496	101203-31-0	688, 735	121788-77-0	329
87413-09-0	494	102029-52-7	857	123333-53-9	446, 722
87720-90-9	192	102525-11-1	690	123333-55-1	412
87720-91-0	192	102783-18-6	426	123333-68-6	381
87749-50-6	699, 700, 720	103128-76-3	758	123334-16-7	383
88020-97-7	193	103225-02-1	467	123334-21-4	697, 743
88020-99-9	365	104371-20-2	366	123334-29-2	633
88496-88-2	505	104371-21-3	366	125829-24-5	633
88929-35-5	890	104499-08-3	759	126050-88-3	809
89059-06-3	481	104706-47-0	449	126456-43-7	265
89976-15-8	452	104713-03-3	684	126456-43-7	715, 717, 718
91159-79-4	271	104714-94-5	692	126949-63-1	687
91510-62-2	824	104830-06-0	404	127184-05-8	429
91682-96-1	898	105308-41-6	846	127280-17-5	451
91742-16-4	731	106018-85-3	544	127382-91-6	633, 670, 936
91742-17-5	731	106268-95-5	443	128625-52-5	722, 724
91742-18-6	731	106868-21-7	636	129476-47-7	632
91742-19-7	731	107054-29-5	428	129880-73-5	652
91742-20-0	730	107439-61-2	528	131042-77-8	723
91742-21-1	730	107492-82-0	239	131274-22-1	735, 750
92212-70-9	630	108347-97-3	898	131347-76-7	778
92307-52-3	782	108448-77-7	226	131380-80-8	677, 727
92484-61-2	132	108503-47-5	661	131380-90-0	728
93634-87-8	505	108703-35-1	647	131380-93-3	677, 727, 728
93670-18-9	291	108750-13-6	631	132063-05-9	816
94041-16-4	725	109276-34-8	864	132071-87-5	684
94041-18-6	725	109669-56-9	302	132873-57-5	362
94051-08-8	897	109682-26-0	260	133882-98-1	880
94218-72-1	808	109704-53-2	535	134029-92-8	939
94348-92-2	646	110199-16-1	759	135105-52-1	911
94548-75-1	512, 637	110827-07-1	260	135113-15-4	913
94594-90-8	226	111770-91-3	462	135113-16-5	913
94790-37-1	410	112022-81-8	717	135616-36-3	712, 724
95197-95-8	53	112022-83-0	717	135616-40-9	712, 724
95456-31-8	516, 736	112068-01-6	436	135620-04-1	675
95530-58-8	759	112926-00-8	606	136030-00-7	265
95977-28-9	897	112945-52-5	606	136030-00-7	714, 717, 718
96476-01-6	627	113303-27-8	239	136316-32-0	913
96638-28-7	813	113303-28-9	239	138124-32-0	675
96638-29-8	813	113303-35-8	239	138517-61-0	733
97731-02-7	786	113350-86-0	239	138800-17-6	547, 548
97945-19-2	223	113975-23-8	404	140472-69-1	290
98327-87-8	714, 725	114200-15-6	290	141556-42-5	731
98977-36-7	413, 414	114234-09-2	503	141556-45-8	696, 731
99685-96-8	909, 911, 916	114269-11-3	660	143314-16-3	756
100243-39-8	449	114615-82-6	656, 708, 939	143314-17-4	756
100312-29-6	473	115383-22-7	909, 912	145224-92-4	651
100439-84-7	660	115459-65-9	443	145626-87-5	118
100622-34-2	500	115514-67-5	299	146651-75-4	286
100646-51-3	896	115826-95-4	672, 675	148263-58-5	812
100646-52-4	896	116797-51-5	783	148461-14-7	740
100760-04-1	290	118336-86-0	419	148461-15-6	740
100760-08-5	896	119110-37-1	642	148672-09-7	850
100760-10-9	896	119189-23-0	809	148719-90-8	436

150493-29-1	<b>914, 916</b>	216299-76-2	<b>756</b>	423118-47-2	<b>524</b>
150760-95-5	<b>166</b>	217189-40-7	<b>675</b>	438553-44-7	<b>525</b>
151120-25-1	<b>636</b>	218290-24-5	<b>734</b>	442905-33-1	<b>707, 749</b>
153218-95-2	<b>914</b>	219143-92-7	<b>675</b>	443347-10-2	<b>749</b>
155116-19-1	<b>916</b>	221012-82-4	<b>707, 749</b>	445473-58-5	<b>756</b>
155155-73-0	<b>435</b>	222400-02-4	<b>743</b>	448245-52-1	<b>756</b>
155371-19-0	<b>756</b>	222400-03-5	<b>680</b>	463941-07-3	<b>717</b>
155679-98-4	<b>914, 916</b>	223906-38-5	<b>758</b>	477873-22-6	<b>245</b>
156150-40-2	<b>653</b>	224311-51-7	<b>678, 722</b>	552331-00-7	<b>404</b>
156397-47-6	<b>709</b>	226909-63-3	<b>917</b>	556907-19-8	<b>590, 687</b>
157437-38-2	<b>422</b>	236388-46-8	<b>525</b>	561328-67-4	<b>524</b>
161265-03-8	<b>673, 710, 751</b>	244187-81-3	<b>729</b>	561328-69-6	<b>526</b>
161337-67-3	<b>693</b>	244229-54-7	<b>916</b>	564483-18-7	<b>679, 738</b>
162157-03-1	<b>741</b>	244237-40-9	<b>917</b>	578743-87-0	<b>682</b>
163169-10-6	<b>700</b>	246047-72-3	<b>706</b>	581093-92-7	<b>681</b>
163169-29-7	<b>700</b>	247940-06-3	<b>722</b>	612065-00-6	<b>681</b>
164858-78-0	<b>740</b>	250285-32-6	<b>683, 723, 729</b>	652157-52-3	<b>871</b>
164931-83-3	<b>675</b>	250611-13-3	<b>675</b>	657408-07-6	<b>737</b>
166328-09-2	<b>523</b>	252726-24-2	<b>523</b>	658080-04-7	<b>918</b>
166328-14-9	<b>524</b>	252976-51-5	<b>694</b>	663199-28-8	<b>756</b>
167171-03-1	<b>740</b>	255728-57-5	<b>736</b>	668984-08-5	<b>526</b>
168297-86-7	<b>759</b>	258873-45-9	<b>692, 742</b>	698999-65-4	<b>693</b>
169689-05-8	<b>733</b>	259860-08-7	<b>415</b>	740816-14-2	<b>716</b>
172138-95-3	<b>734</b>	261948-85-0	<b>684</b>	744998-97-8	<b>563, 628</b>
172152-19-1	<b>436, 719</b>	262297-13-2	<b>756</b>	755017-61-9	<b>289</b>
172222-30-9	<b>705</b>	267006-25-7	<b>524</b>	797757-81-4	<b>728</b>
173593-25-4	<b>736</b>	284049-75-8	<b>755</b>	805315-08-6	<b>682</b>
174501-64-5	<b>756</b>	286014-42-4	<b>723</b>	805315-09-7	<b>681</b>
174501-65-6	<b>756</b>	287105-48-0	<b>695</b>	811862-77-8	<b>674</b>
174645-81-9	<b>756</b>	287410-78-0	<b>697</b>	849592-74-1	<b>681</b>
174883-18-2	<b>809</b>	287410-79-1	<b>698</b>	849921-25-1	<b>701</b>
174899-66-2	<b>756</b>	300344-02-9	<b>687</b>	851609-33-1	<b>724</b>
174899-82-2	<b>756</b>	303136-74-5	<b>455</b>	851770-14-4	<b>724</b>
174899-83-3	<b>755</b>	307926-51-8	<b>675</b>	851942-89-7	<b>753</b>
175833-78-0	<b>913</b>	308068-56-6	<b>924</b>	852445-81-9	<b>678</b>
176762-62-5	<b>675</b>	312619-45-7	<b>853</b>	852445-82-0	<b>678</b>
181486-37-3	<b>220</b>	312693-53-1	<b>404</b>	854045-93-5	<b>678</b>
182627-81-2	<b>743</b>	312693-83-7	<b>532, 855</b>	854045-94-6	<b>679</b>
188264-84-8	<b>675</b>	312959-24-3	<b>746</b>	862095-79-2	<b>714, 716</b>
192463-40-4	<b>676</b>	313223-04-0	<b>844</b>	862095-91-8	<b>716</b>
192863-36-8	<b>525</b>	325168-88-5	<b>677</b>	865812-10-8	<b>717</b>
192863-37-9	<b>526</b>	325168-89-6	<b>677</b>	866395-16-6	<b>676, 708</b>
199588-80-2	<b>752</b>	329735-68-4	<b>688, 735</b>	866641-66-9	<b>678</b>
205815-80-1	<b>331</b>	329735-77-5	<b>687</b>	869858-84-4	<b>756</b>
205815-80-1	<b>687</b>	329976-74-1	<b>526</b>	871231-40-2	<b>526</b>
205815-80-1	<b>736</b>	333380-86-2	<b>672, 749</b>	871231-41-3	<b>524</b>
205927-03-3	<b>688, 735</b>	336096-71-0	<b>811</b>	886059-84-3	<b>735</b>
206752-36-5	<b>272</b>	342789-81-5	<b>756</b>	886061-48-9	<b>683</b>
207121-53-7	<b>871</b>	345909-26-4	<b>532, 855</b>	887919-35-9	<b>674</b>
207124-65-0	<b>703</b>	356066-52-9	<b>914</b>	887938-70-7	<b>753</b>
207233-95-2	<b>530</b>	362524-23-0	<b>707, 749</b>	903571-02-8	<b>715</b>
207300-70-7	<b>826</b>	364732-88-7	<b>676</b>	929270-13-3	<b>728</b>
207399-11-9	<b>585</b>	366491-15-8	<b>756</b>	943552-04-3	<b>731</b>
207683-22-5	<b>532</b>	370865-89-7	<b>756</b>	1030390-38-5	<b>245</b>
207737-97-1	<b>855</b>	379669-72-4	<b>504</b>	1104636-68-1	<b>732</b>
212133-11-4	<b>700</b>	401788-98-5	<b>756</b>	1104636-73-8	<b>750</b>
212210-73-6	<b>701</b>	419581-64-9	<b>697, 698</b>	1105069-27-9	<b>732</b>

1145982-32-6	<i>698</i>
1217463-35-8	<i>759</i>
1260376-31-5	<i>913</i>

COOPER UNION LIBRARY



3 1206 04669543 9

WITHDRAWN FROM  
COOPER UNION LIBRARY

Wilfred L. F. Armarego • Christina Li Lin Chai

# PURIFICATION OF LABORATORY CHEMICALS

SEVENTH EDITION

- Improves efficiency, results and safety by providing critical information for day-to-day lab and processing work
- Searchable electronic database allows instant access to accurate, reliable information on processes and techniques of purification along with detailed physical properties
- New commercially available chemical products and processing techniques are covered; hazards and safety remain central to the book

A bestseller since 1966, Purification of Laboratory Chemicals keeps engineers, scientists, chemists, biochemists and students up-to-date with the processes for purification of the chemical reagents with which they work, and guides readers on critical safety and hazards for the safe handling of chemicals and processes. The 7th Edition is updated and provides expanded coverage of the latest chemical products and processing techniques, safety and hazards. The book has been reorganized, has a new chapter on Catalysts and one on Nanomaterials, and is fully indexed by CAS Registry Numbers. Compounds are now grouped to make navigation easier, and extensive literature references for all substances and techniques have been added.

#### About the authors

W.L.F. Armarego

Department of Structural Biology, The John Curtin School of Medical Research, Australian National University, Canberra

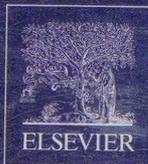
Wilfred L. F. Armarego graduated BSc (Special, Hons) in 1953 and PhD from the University of London in 1956 and came to Australia in that year. After two years at the Central Research Laboratories (ICIANZ) in Melbourne, where he worked on plant growth substances, and one year on potentially carcinogenic polycyclic aromatic hydrocarbons at the University of Melbourne as Senior Demonstrator in Organic Chemistry, he joined the Department of Medical Chemistry as a Research Fellow in 1960. He became a Fellow in 1963 and was awarded a DSc degree (London) in 1968. He was promoted to Senior Fellow in 1967 and began research work on the biochemistry and molecular biology of pteridine-requiring enzymes related to the inherited metabolic disease phenylketonuria and its variants. He was head of the Protein Biochemistry Group and Pteridine Biochemistry Laboratory. In addition to writing more than 100 original research papers and reviews in heterocyclic chemistry, enzymology and molecular cloning, he has written three other books as sole author. He is now a Visiting Fellow in the John Curtin School of Medical Research, and member of the editorial boards of 'Medicinal Research Reviews' and 'Pteridines' journals. He is a Chartered Chemist, a Fellow of the Royal Society of Chemistry and a Fellow of the Royal Australian Chemical Institute.

Christina L. L. Chai

Department of Pharmacy, National University of Singapore & Institute of Chemical and Engineering Sciences, Agency for Science Technology and Research (A\*STAR), Singapore

Christina Li Lin Chai obtained her BSc (Hons) from the University of Canterbury, Christchurch, New Zealand and her PhD in synthetic organic chemistry from the Research School of Chemistry, Australian National University, Canberra.

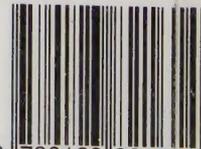
She joined the Institute of Chemical and Engineering Sciences as a Principal Scientist and Programme Manager of the New Synthesis Techniques and Applications (NSTA) Programme from 2004 till 2011. Since August 2011, Dr Chai holds a joint appointment as Associate Professor in the Department of Pharmacy, National University of Singapore and the Institute of Chemical and Engineering Sciences, A\*STAR. Dr Chai's professional achievements include the DAAD Fellowship and Alexander von Humboldt Fellowship and she is a member of the American Chemical Society, and the Singapore Institute of Chemistry. Her research interests are in the areas of medicinal chemistry, green technologies and synthesis of bioactive compounds and elucidation of their mechanisms.



Butterworth-Heinemann

An imprint of Elsevier  
store.elsevier.com

ISBN 978-0-12-382161-4



9 780123 821614