

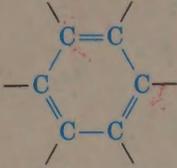
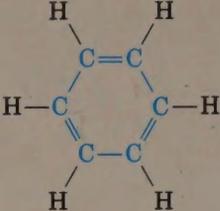
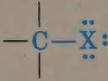
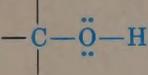
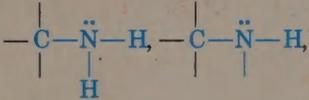
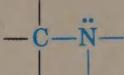
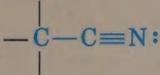
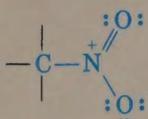
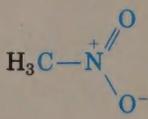
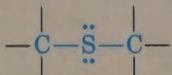
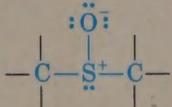
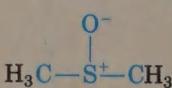
# McMURRY

3

ORGANIC CHEMISTRY



# Some Functional Groups

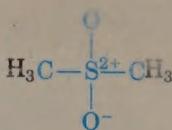
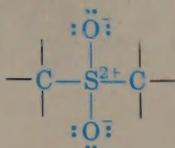
Family name	Functional group structure <sup>a</sup>	Simple example	Name ending
Alkane	(Contains only C—H and C—C single bonds)	CH <sub>3</sub> CH <sub>3</sub>	-ane Ethane
Alkene		H <sub>2</sub> C=CH <sub>2</sub>	-ene Ethene (Ethylene)
Alkyne	—C≡C—	H—C≡C—H	-yne Ethyne (Acetylene)
Arene			None Benzene
Halide	 (X = F, Cl, Br, I)	H <sub>3</sub> C—Cl	None Chloromethane
Alcohol		H <sub>3</sub> C—O—H	-ol Methanol
Ether		H <sub>3</sub> C—O—CH <sub>3</sub>	ether Dimethyl ether
Amine	 	H <sub>3</sub> C—NH <sub>2</sub>	-amine Methylamine
Nitrile		H <sub>3</sub> C—C≡N	-nitrile Ethanenitrile (Acetonitrile)
Nitro			None Nitromethane
Sulfide		H <sub>3</sub> C—S—CH <sub>3</sub>	sulfide Dimethyl sulfide
Sulfoxide			sulfoxide Dimethyl sulfoxide

<sup>a</sup>The bonds whose connections aren't specified are assumed to be attached to carbon or hydrogen atoms in the rest of the molecule.

Family name

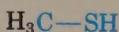
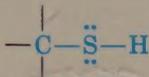
Name ending

Sulfone



*sulfone*  
Dimethyl sulfone

Thiol

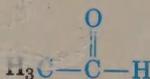
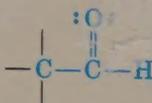


*-thiol*  
Methanethiol

Date

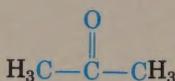
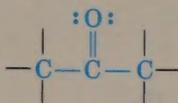
Carbonyl,  $\text{---C---}$

Aldehyde



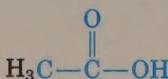
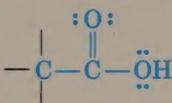
*-al*  
Ethanal  
(Acetaldehyde)

Ketone



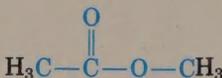
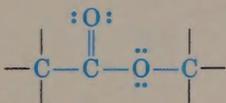
*-one*  
Propanone  
(Acetone)

Carboxylic acid



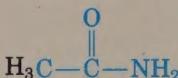
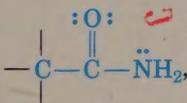
*-oic acid*  
Ethanoic acid  
(Acetic acid)

Ester

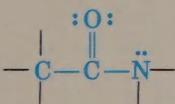
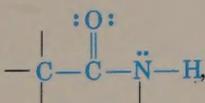


*-oate*  
Methyl ethanoate  
(Methyl acetate)

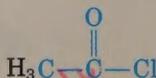
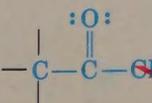
Amide



*-amide*  
Ethanamide  
(Acetamide)

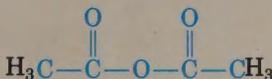
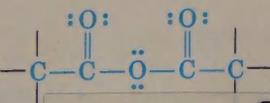


Carboxylic acid  
chloride



*-oyl chloride*  
Ethanoyl chloride  
(Acetyl chloride)

Carboxylic acid  
anhydride



*-oic anhydride*  
Ethanoic anhydride  
(Acetic anhydride)

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# Organic Chemistry

Third Edition

**John McMurry**

Cornell University



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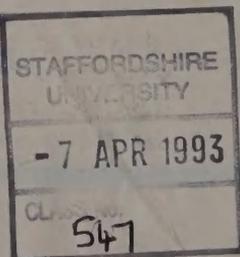
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# Preface

In the first edition of this text, I stated that my goal was to write a readable and effective *teaching* text, while providing an accurate and up-to-date view of organic chemistry. The outstanding reception of the two previous editions by both students and faculty suggests that the goal has been met. This third edition of *Organic Chemistry* builds on the strengths of the first two. The features that made those editions successful have been improved and new ones have been added:

- The writing, already lucid and accessible to students, has again been reworked at the sentence level. Particular attention has been paid to such traditionally difficult subjects as stereochemistry and nucleophilic substitutions, both of which topics have been recast.
- The artwork, already attractive, has been completely redone with use of full color to highlight chemical transformations, the addition of many computer-generated structures to aid in three-dimensional perception, and a greater use of graphics for data presentation.
- The problem sets, already numerous, have been expanded even further. More than 100 new problems have been added, many of them at a challenging level.
- Still other changes in this third edition include to addition of more solved practice problems, a more thorough review of fundamental topics from general chemistry, an expanded treatment of acid/base chemistry, the addition of new topics such as the Birch reduction, the deletion of older topics such as a discussion of tetrahydropyranyl protecting groups, an updating of synthetic methods such as the replacement of *m*-chloroperoxybenzoic acid by magnesium monoperoxyphthalate for alkene epoxidations, an appendix giving short answers to most in-text problems, and a thorough updating of nucleic acid chemistry.

## Organization

This book uses a dual organization that blends the traditional functional-group approach with reaction mechanisms. The primary organization is by functional group, beginning with the simple (alkenes) and progressing to the more complex. Within this primary organization, however, heavy

emphasis is placed on explaining the fundamental mechanistic similarities of reactions. This is particularly evident in the chapters on carbonyl-group chemistry (Chapters 19–23) where mechanistically related reactions like the aldol and Claisen condensations are treated together.

Insofar as possible, topics have been arranged in a modular way. The chapters on simple hydrocarbons are grouped together (Chapters 3–8), the chapters on spectroscopy are grouped together (Chapters 12–14), and the chapters on carbonyl-group chemistry are grouped together (Chapters 19–23). This organization brings to these subjects a cohesiveness not found in other texts and allows the instructor the flexibility to teach in an order different from that presented in this book.

### The Lead-Off Reaction: Addition of HBr to Alkenes

Many students attach great importance to a text's lead-off reaction, because it is the first reaction they see and is discussed in such detail. This book uses a simple polar reaction, the addition of HBr to an alkene, as the lead-off to illustrate the general principles of organic reactions. This choice has the advantage of relative simplicity (no prior knowledge of chirality or kinetics is required) yet it is an important polar reaction on a common functional group. As such, I believe that this choice is a much more useful introduction to functional-group chemistry than a lead-off such as free-radical alkane chlorination.

### Coverage

The coverage in this book is up-to-date, reflecting the important advances of the last decade. For example,  $^{13}\text{C}$  NMR is introduced as a routine spectroscopic tool; the use of lithium diisopropylamide as a base for carbonyl-group alkylations is discussed; selenoxide elimination as a method of introducing double bonds is covered; silyl ether protecting groups are discussed; the chemistry of DNA sequencing and synthesis is treated in detail; and entire chapters are devoted to polymer chemistry and to pericyclic reactions.

### Organic Synthesis

Organic synthesis is emphasized as a teaching device to help students learn to organize and work with the large body of factual information that makes up organic chemistry. Two sections, the first in Chapter 8 (Alkynes) and the second in Chapter 16 (Aromatics), clearly explain the thought processes involved in working synthesis problems. The value of starting from what is known and logically working backwards one step at a time is emphasized. Also important in this respect are several summaries included in the accompanying *Study Guide and Solutions Manual*. A summary of functional-group

reactions, another on functional group preparations, and a third on reagents used in organic synthesis are enormously useful when working problems.

## Pedagogy

Every effort has been made to make this third edition as effective, clear, and readable as possible—to make it easy to learn from. In addition to features already mentioned, a wide assortment of additional pedagogical devices are used:

Paragraphs start with summary sentences.

Transitions between paragraphs and between topics are smooth.

New concepts are introduced only as needed and are immediately illustrated with concrete examples.

Extensive use is made of computer-generated three-dimensional art and carefully rendered stereochemical formulas.

Extensive cross-referencing to earlier material is used.

Numerous summaries are included, both within chapters and at the ends of chapters.

More than 1500 problems of varying difficulty are included within the text and at the ends of chapters.

Use of an innovative vertical format to explain reaction mechanisms, introduced so successfully in the first edition, is expanded. Mechanisms are printed vertically, while explanations of the changes taking place in each step are printed next to the reaction arrow. This format allows the reader to see easily what is occurring at each step in a reaction without having to jump back and forth between the text and structures.

## Appendixes

The back of this book contains a wealth of material helpful for learning organic chemistry. Included are a large glossary of more than 250 terms, an explanation of how to name polyfunctional organic compounds, and answers to most in-text problems.

## Study Guide and Solutions Manual

A carefully prepared *Study Guide and Solutions Manual* accompanies this text. Written by Susan McMurry, this companion volume answers all in-text and end-of-chapter problems and explains in detail how answers are obtained. In addition, the following supplemental materials are included: a list of study goals for each chapter, a list of suggested readings, a summary of name reactions, a summary of methods for preparing functional groups, a summary of functional-group reactions, a summary of the uses of important reagents, and tables of spectroscopic information.

## Acknowledgments

It is a great pleasure to thank the many people whose help and suggestions were so valuable in preparing this third edition. Foremost is my wife, Susan, who read, criticized, and improved the manuscript, and who was my constant companion throughout all stages of this book's development. Among the reviewers providing valuable comments were John Belletire, University of Cincinnati; Kenneth S. Feldman, Pennsylvania State University; Martin Feldman, Howard University; John Henderson, Jackson Community College; Stephen Hixon, University of Massachusetts; Andrew S. Kende, University of Rochester; John T. Landrum, Florida International University; Kenneth L. Marsi, California State University—Long Beach; Eugene A. Mash, University of Arizona; William E. Russey, Juniata College; John R. Stille, Michigan State University.

Of course, the advice of the reviewers of the previous editions continues to be invaluable. For the first edition: Robert Benkeser (Purdue University), Donald E. Bergstrom (University of North Dakota), Weston T. Borden (University of Washington), Larry Bray (Miami Dade Community College), William D. Closson (State University of New York, Albany), Paul L. Cook (Albion College), Otis Dermer, (Oklahoma State University), Linda Domel-Smith, David Harpp (McGill University), David Hart (Ohio State University), Norbert Hepfinger (Rensselaer Polytechnic Institute), Werner Herz (Florida State University), Paul R. Jones (North Texas State University), Thomas Katz (Columbia University), Paul E. Klinedinst, Jr. (California State University, Northridge), James G. Macmillan (University of Northern Iowa), Monroe Moosnick (Transylvania University), Harry Morrison (Purdue University), Cary Morrow (University of New Mexico), Wesley A. Pearson (St. Olaf College), Frank P. Robinson (University of Victoria), Neil E. Schore (University of California, Davis), Gerald Selter (California State University, San Jose), Ernest Simpson (California State Polytechnic University, Pomona), Walter Trahanovsky (Iowa State University), Harry Ungar (Cabrillo College), Joseph J. Villafranca (Pennsylvania State University), Daniel P. Weeks (Northwestern University), Walter Zajac (Villanova University), and Vera Zalkow (Kennesaw College). For the second edition: Ronald Caple (University of Minnesota, Duluth), John Cawley (Villanova University), Clair Cheer (University of Rhode Island), George Clemons (Bowling Green State University), Dennis Davis (New Mexico State University), John Hogg (Texas A&M University), Paul Hopkins (University of Washington), John Huffman (Clemson University), Glen Kauffman (East Mennonite College), Joseph Lambert (Northwestern University), Thomas Livinghouse (University of Montana), Clarence Murphy (East Stroudsburg University), Oliver Muscio (Murray State University), Russell Petter (University of Pittsburgh), Neil Potter (Susquehanna University), Naser Pourahmady (Southwest Missouri State University), William Russey (Juniata College), Jan Simek (California Polytechnic State University), Marcus W. Thomsen (Franklin & Marshall College), Daniel Weeks (Northwestern University), and David Wiemer (University of Iowa).

John McMurry  
Cornell University

# A Note for Students

We have similar goals: Yours is to learn organic chemistry, and mine is to do everything possible to help you learn. It's going to require some work on your part, but the following hints should prove helpful:

*Don't read the text immediately.* As you begin each new chapter, look it over first. Read the introductory paragraphs, find out what topics will be covered, and then turn to the end of the chapter and read the summary. You'll be in a much better position to understand new material if you first have a general idea of where you're going.

*Work the problems.* There are no shortcuts here; working problems is the only way to learn organic chemistry. The practice problems show you how to approach the material, the in-text problems at the ends of most sections provide immediate practice, and the end-of-chapter problems provide both additional drill as well as some real challenges. Full answers and explanations for all problems are given in the accompanying *Study Guide and Solutions Manual*.

*Use the Study Guide.* The *Study Guide and Solutions Manual* that accompanies this text gives complete solutions to all problems, an outline and list of study goals for each chapter, tables of supplementary material, and additional unsolved problems.

*Use the Appendixes.* The back of this book and the back of the *Study Guide* contain a wealth of supplementary material, including a large glossary of terms, an explanation of how to name polyfunctional organic compounds, and extensive lists summarizing the reagents commonly used in organic chemistry, the ways in which different functional groups are prepared, and the reactions that functional groups undergo. This material can be extremely useful when you're studying for an exam and for getting an overview of organic chemistry. Find out what's there now, so you'll know where to find it when you need help.

*Ask questions.* Faculty members and teaching assistants are there to help you. Most of them will turn out to be genuinely nice people with a sincere interest in helping you learn.

*Use molecular models.* Organic chemistry is a three-dimensional science. Although this book uses many careful drawings to help you visualize molecules, there's no substitute for building a molecular model and turning it around in your hands.

Good luck. I sincerely hope you enjoy learning organic chemistry and come to see the beauty and logic of its structure. I heard from many students who used the first two editions of this book and would be glad to receive more comments and suggestions from those who use this new edition.

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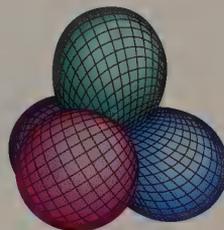
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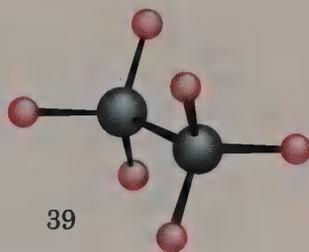
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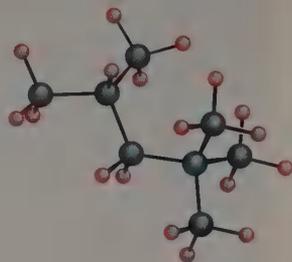
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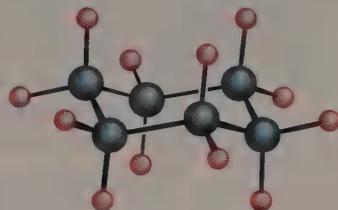
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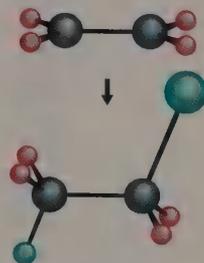
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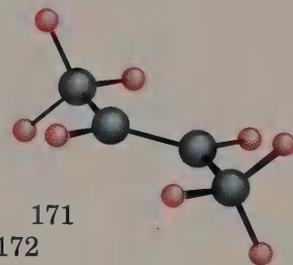


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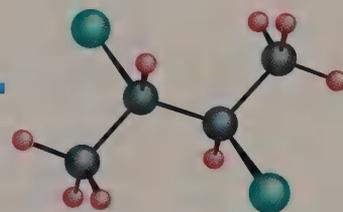
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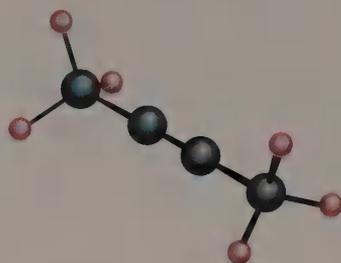


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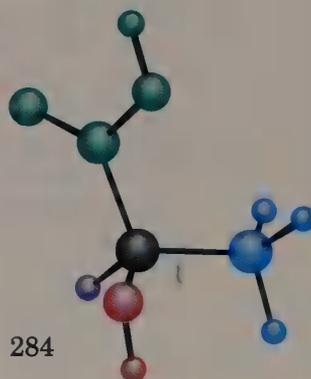
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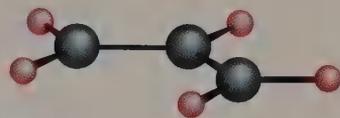


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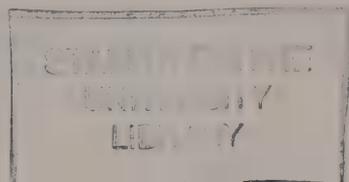
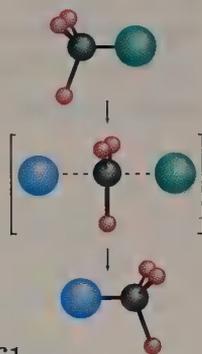
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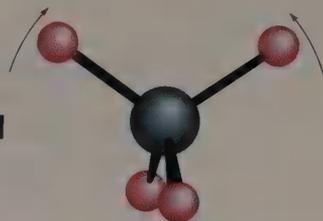
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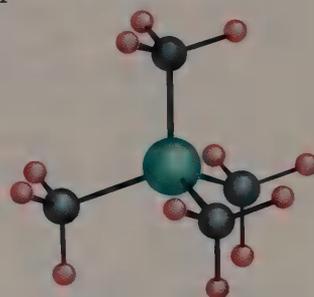
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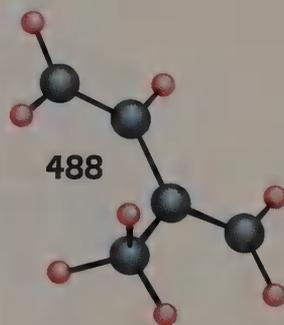
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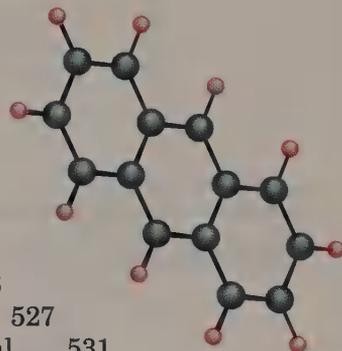


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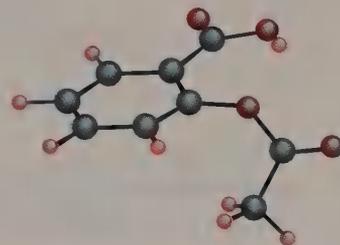
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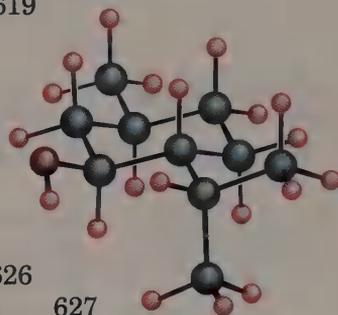
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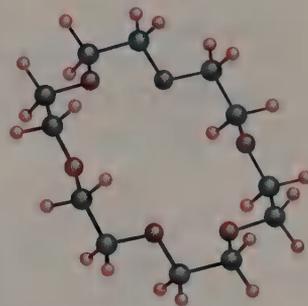


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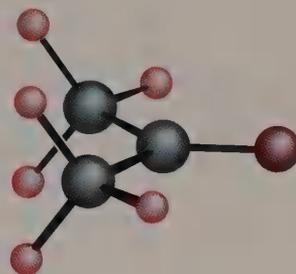
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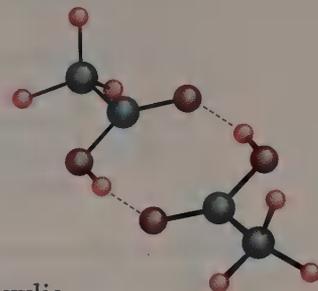
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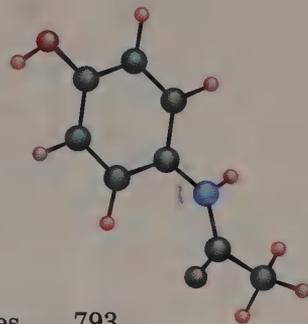
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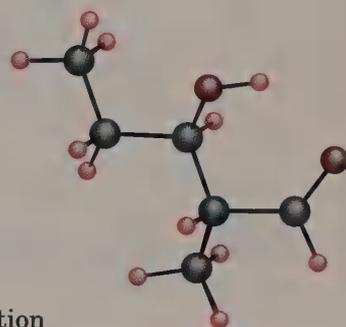
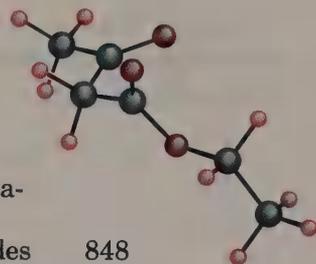
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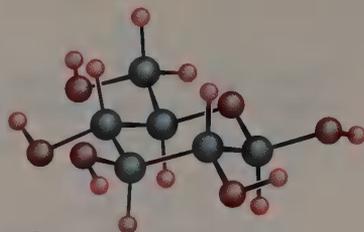
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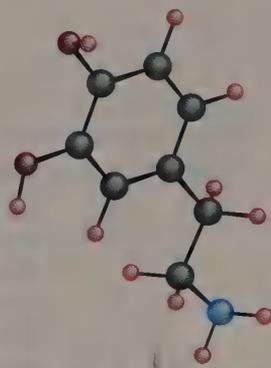
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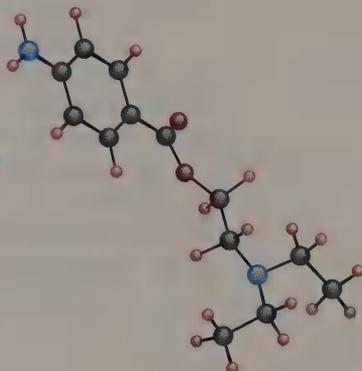
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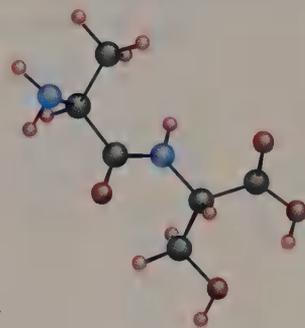


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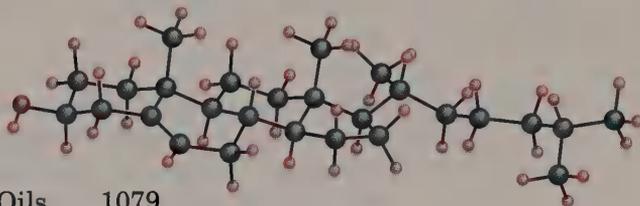
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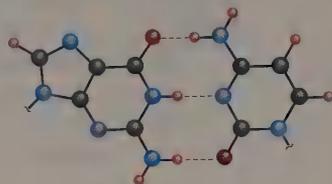


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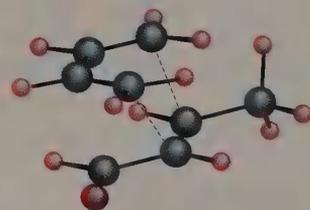
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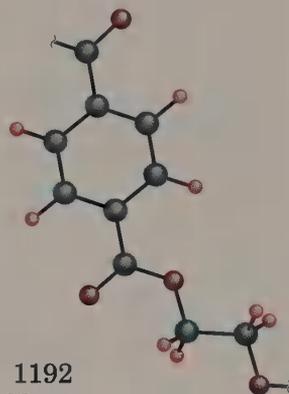


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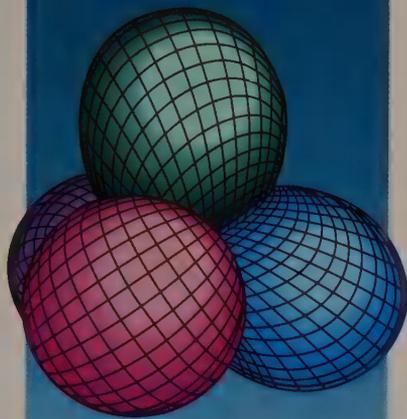


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# Chapter 1

## Structure and Bonding

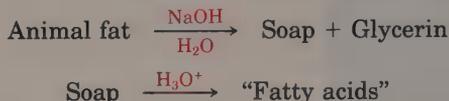
What is organic chemistry, and why should you study it? The answers are all around you. Every living organism is made of organic chemicals. The proteins that make up our hair, skin, and muscles; the DNA that controls our genetic heritage; the foods we eat; the clothes we wear; and the medicines we take are all organic chemicals. Anyone with a curiosity about life and living things must have a fundamental understanding of organic chemistry.

The foundations of organic chemistry were built in the mid-eighteenth century when chemistry was evolving from an alchemist's art into a modern science. At that time, unexplainable differences were noted between substances obtained from living sources and those from minerals. Compounds from plants and animals were often difficult to isolate and purify. Even when pure, they were often difficult to work with and tended to decompose more easily than compounds from minerals. The Swedish chemist Torbern Bergman was the first to express this difference between "organic" and "inorganic" substances in 1770, and the term *organic chemistry* soon came to mean the chemistry of compounds from living organisms.

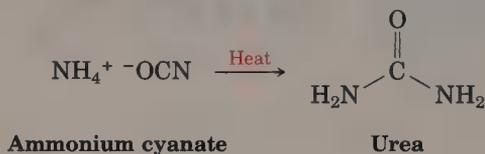
To many chemists at the time, the only explanation for the difference in behavior between organic and inorganic compounds was that organic compounds contained a peculiar "vital force" as a result of their origin in living organisms. One consequence of this vital force, chemists believed, was that organic compounds couldn't be prepared and manipulated in the laboratory as could inorganic compounds.

Although the vitalistic theory was believed by many influential chemists, its acceptance was by no means universal, and it's doubtful that the development of organic chemistry was much delayed. As early as 1816, the

theory received a heavy blow when Michel Chevreul<sup>1</sup> found that soap, prepared by the reaction of alkali with animal fat, could be separated into several pure organic compounds, which he termed “fatty acids.” Thus, for the first time, one organic substance (fat) had been converted into others (fatty acids plus glycerin) without the intervention of an outside vital force.<sup>2</sup>



A little more than a decade later, the vitalistic theory suffered still further when Friedrich Wöhler<sup>3</sup> discovered in 1828 that it was possible to convert the “inorganic” salt, ammonium cyanate, into the previously known “organic” substance, urea.



By the mid-nineteenth century, the weight of evidence was clearly against the vitalistic theory, and William Brande<sup>4</sup> wrote in 1848 that “no definite line can be drawn between organic and inorganic chemistry. . . . Any distinctions . . . must for the present be merely considered as matters of practical convenience calculated to further the progress of students.”

Chemistry today is unified. The same principles that explain the simplest inorganic compounds also explain the most complex organic ones. The only distinguishing characteristic of organic chemicals is that all contain the element carbon. Nevertheless, the division between organic and inorganic chemistry, which began for historical reasons, maintains its “practical convenience . . . to further the progress of students.”

**Organic chemistry**, then, is the study of carbon compounds. Carbon, atomic number 6, is a second-row element whose position in the periodic table is shown in Figure 1.1. Although carbon is the principal element in organic compounds, most also contain hydrogen, and many contain nitrogen, oxygen, phosphorus, sulfur, chlorine, and other elements.

Why is carbon special? What is it that sets carbon apart from all other elements in the periodic table? The answers to these questions derive from

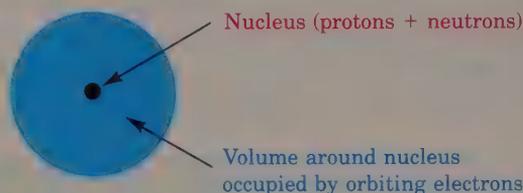
<sup>1</sup>Michel Eugène Chevreul (1786–1889); b. Angers, France; educated at Paris, Muséum d’Histoire Naturelle; professor of physics, Lycée Charlemagne (1813); professor of chemistry (1830).

<sup>2</sup>In the equations that follow, a single arrow,  $\rightarrow$ , is used to indicate an actual reaction. Later in this book you’ll also see forward and backward arrows,  $\rightleftharpoons$ , indicating equilibrium, and a double arrow,  $\rightleftarrows$ , indicating a multistep transformation whose individual steps aren’t specified.

<sup>3</sup>Friedrich Wöhler (1800–1882); b. Eschersheim; studied at Heidelberg (Gmelin); professor, Göttingen (1836–1882).

<sup>4</sup>William Thomas Brande (1788–1866); b. London; lecturer in chemistry, London (1808); Royal Institution (1813–1854).





**Figure 1.2** A schematic view of an atom. The dense, positively charged nucleus contains most of the atom's mass and is surrounded by negatively charged electrons.

$10^{-10}$  m.<sup>5</sup> To give you an idea of how small this is, a thin pencil line is about 3 million carbon atoms across.

A specific atom is described by its **atomic number**, which is the number of protons in the atom's nucleus, and its **mass number**, which is the total of protons plus neutrons. All the atoms of a given element have the same atomic number—1 for hydrogen, 6 for carbon, 17 for chlorine, and so on—but they can have different mass numbers depending on how many neutrons they contain. The *average* mass number of a great many atoms of an element is called the element's **atomic weight**. Because it's an average, it's usually not an integer (1.008 for hydrogen, 35.453 for chlorine, and so on).

## 1.2 The Structure of Atoms: Orbitals

How are the electrons distributed in an atom? A major breakthrough in our understanding of atomic structure occurred in 1926 when the theory of **quantum mechanics** was proposed independently by Paul Dirac, Werner Heisenberg, and Erwin Schrödinger.<sup>6</sup> All three formulations are mathematical expressions for describing the electronic structure of atoms, but Schrödinger's is the one most commonly used by chemists.

The Schrödinger equation says that the motion of an electron around the nucleus can be described mathematically by what is known as a **wave equation**—the same kind of expression that's used to describe the motion of waves in a fluid. The solution to a wave equation is called a **wave function**. If we could determine the wave function for every electron in an atom,

<sup>5</sup>The angstrom is still used by most organic chemists, even though it has been replaced in SI (Système International) units by the picometer (pm), where 1 pm =  $10^{-12}$  m and 1 Å = 100 pm. Because of the easy decimal conversion, only angstrom measurements will be given in this book.

<sup>6</sup>Erwin Schrödinger (1887–1961); b. Vienna, Austria; University of Vienna (1910); assistant, University of Vienna (1910); assistant to Max Wein, University of Stuttgart, Germany (1920); professor of physics, the universities of Zurich, Berlin, Graz, and Dublin; Nobel Prize in physics (1933).

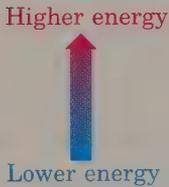
we would have a complete electronic description of that atom. In practice, wave equations are mathematically so complex that only approximate solutions can be obtained, even with the fastest computers now available. These approximate solutions agree so well with experimental facts, though, that quantum mechanics is a universally accepted theory for understanding atomic structure.

How can we interpret quantum mechanical wave functions in terms of physical reality? A good way of viewing a wave function is to think of it as an expression that predicts the volume of space around a nucleus where an electron can be found. Though we can never know the exact position of an electron at a given moment, the wave function tells us where we would be most likely to find it.

According to quantum mechanics, electrons can be thought of as belonging to different layers, or **shells**, at various distances from the nucleus. The farther a shell is from the nucleus, the more electrons it can hold and the greater the energies of those electrons. Thus, an atom's lowest-energy electrons occupy the first shell, which is nearest the nucleus and has a capacity of only two electrons. The second shell is farther from the nucleus and can hold eight electrons; the third shell is still farther from the nucleus and can hold eighteen electrons; and so on, as shown in Table 1.1.

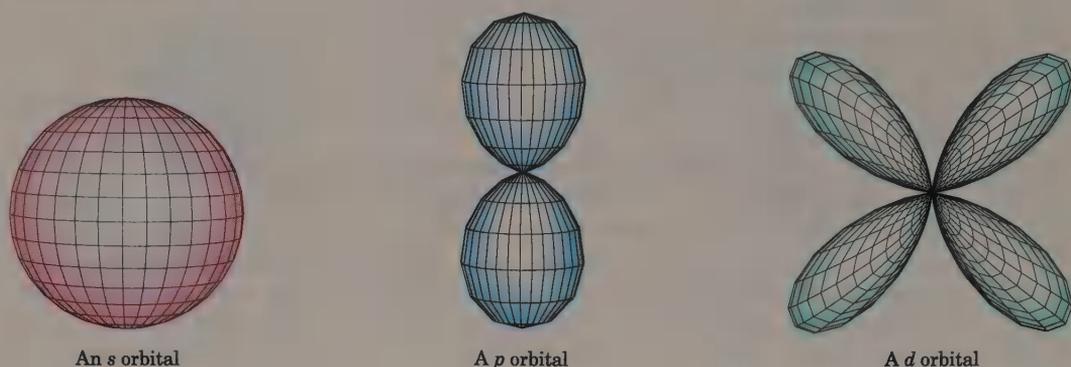
**TABLE 1.1** Distribution of Electrons into Shells

<i>Number of shell</i>	<i>Electron capacity of shell</i>
Fourth	32
Third	18
Second	8
First	2



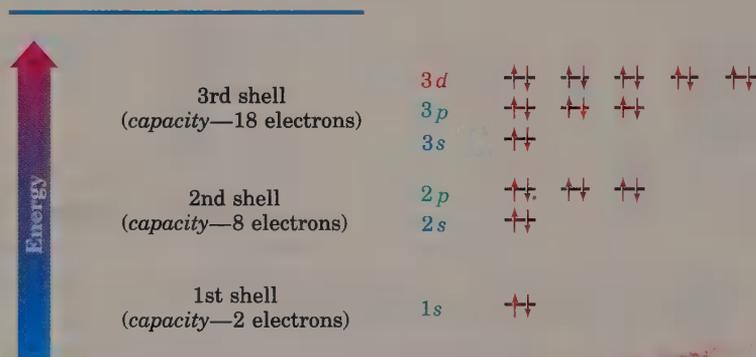
Within each shell, electrons are further grouped by pairs into specific regions of space called **orbitals**. It's helpful to think of an orbital as a kind of time-lapse photograph of an electron's movement around the nucleus. Such a photograph would show the orbital as a blurry cloud indicating where the electron has been. This electron cloud doesn't have a sharp boundary, but for practical purposes we can set the limits by saying that an orbital represents the space where an electron spends most (90–95%) of its time.

What do orbitals look like? The exact shape and size of an electron's orbital depend on its energy level and are described mathematically by its quantum mechanical wave function. There are four kinds of orbitals, denoted *s*, *p*, *d*, and *f*. Of the four, we'll be concerned primarily with *s* and *p* orbitals since these play the most important role in organic chemistry. The *s* orbitals have a spherical shape with the nucleus at the center, the *p* orbitals have a dumbbell shape, and four of the five *d* orbitals have a cloverleaf shape, as shown in Figure 1.3.



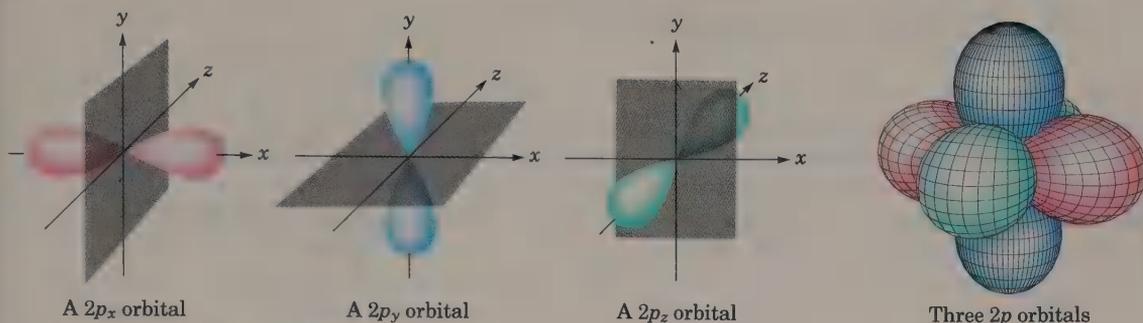
**Figure 1.3** Computer-generated shapes of  $s$ ,  $p$ , and  $d$  orbitals. The  $s$  orbitals are spherical, the  $p$  orbitals are dumbbell shaped, and most  $d$  orbitals are cloverleaf shaped.

Different shells have different numbers and kinds of orbitals, as indicated in Figure 1.4. The lowest-energy electrons occupy the  $1s$  orbital. Next in energy after the  $1s$  electrons are the  $2s$  electrons. Because they're higher in energy,  $2s$  electrons are farther from the positively charged nucleus on average, and their spherical orbital is somewhat larger than that of  $1s$  electrons.



**Figure 1.4** The distribution of electrons in an atom. The first shell holds a maximum of two electrons in one  $1s$  orbital; the second shell holds a maximum of eight electrons in one  $2s$  and three  $2p$  orbitals; the third shell holds a maximum of eighteen electrons in one  $3s$ , three  $3p$ , and five  $3d$  orbitals; and so on. The two electrons in each orbital are represented by up and down arrows  $\uparrow\downarrow$ .

The  $2p$  electrons are next higher in energy. As Figure 1.5 indicates, there are three  $2p$  orbitals that are equal in energy and are oriented in space such that each is perpendicular to the other two. They're denoted  $2p_x$ ,  $2p_y$ , and  $2p_z$  to show on which axis they lie. Note that the plane passing between the two lobes of a  $p$  orbital is a region of zero electron density. Such



**Figure 1.5** Shapes of the  $2p$  orbitals. Each of the three mutually perpendicular dumbbell-shaped orbitals has a nodal plane passing between its two lobes.

a region of zero electron density is called a **node**, and the plane is called a **nodal plane**. Nodes have important consequences with respect to chemical reactivity that we'll take up in Chapter 30.

Still higher in energy are the  $3s$  orbital (spherical),  $3p$  orbitals (dumbbell shaped),  $4s$  orbital (spherical), and  $3d$  orbitals. There are five  $3d$  orbitals of equal energy, but as previously remarked, we won't be too concerned with them. Note, though, that the  $d$  orbital shown in Figure 1.3 has four lobes and two nodal planes.

## 1.3 The Structure of Atoms: Electronic Configurations

The lowest energy arrangement, or **ground-state electronic configuration**, of an atom is a description of the orbitals that the atom's electrons occupy. This arrangement can be found from a knowledge of atomic orbitals and their energy levels by following three rules:

1. **Always** fill the lowest energy orbitals first (called the **aufbau principle**).
2. Only two electrons can be put into each orbital, and they must be of opposite spin<sup>7</sup> (called the **Pauli exclusion principle**).
3. If two or more empty orbitals of equal energy are available, put one electron in each with the spins parallel until all are half-full (called **Hund's rule**).

Let's look at some examples to see how these rules are applied. Hydrogen, the lightest element, has only one electron, which we assign to the

<sup>7</sup>For the purposes of quantum mechanics, electrons are considered to spin around an axis in much the same way that the earth spins. This spin can have two equal and opposite orientations, denoted as up  $\uparrow$  and down  $\downarrow$ .

lowest-energy orbital. This gives hydrogen a  $1s$  ground-state configuration. Carbon has six electrons, and a ground-state configuration  $1s^2 2s^2 2p_x 2p_y$  is arrived at by applying the three rules.<sup>8</sup> These and other examples are shown in Table 1.2.

TABLE 1.2 Ground-State Electronic Configurations of Some Elements

Element	Atomic number	Configuration	Element	Atomic number	Configuration
Hydrogen	1	$1s \uparrow$	Lithium	3	$2s \uparrow$ $1s \uparrow\downarrow$
Carbon	6	$2p \uparrow \uparrow \_$ $2s \uparrow\downarrow$ $1s \uparrow\downarrow$	Neon	10	$2p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$ $2s \uparrow\downarrow$ $1s \uparrow\downarrow$
Sodium	11	$3s \uparrow$ $2p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$ $2s \uparrow\downarrow$ $1s \uparrow\downarrow$	Argon	18	$3p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$ $3s \uparrow\downarrow$ $2p \uparrow\downarrow \uparrow\downarrow \uparrow\downarrow$ $2s \uparrow\downarrow$ $1s \uparrow\downarrow$

PROBLEM .....

- 1.1 How many electrons does each of these elements have in its outermost electron shell?  
 (a) Potassium (b) Aluminum (c) Krypton

PROBLEM .....

- 1.2 Give the ground-state electronic configuration for these elements:  
 (a) Boron (b) Phosphorus (c) Oxygen (d) Chlorine
- .....

## 1.4 Development of Chemical Bonding Theory

By the mid-nineteenth century, the new science of chemistry was developing rapidly and chemists had begun to probe the forces holding molecules together. In 1858, August Kekulé<sup>9</sup> and Archibald Couper<sup>10</sup> independently proposed that, in all organic compounds, carbon has four “affinity units.”

<sup>8</sup>A superscript is used here to represent the number of electrons at a particular energy level. For example,  $1s^2$  indicates that there are two electrons in the  $1s$  orbital. No superscript is used when there is only one electron in an orbital.

<sup>9</sup>Friedrich August Kekulé (1829–1896); b. Darmstadt; University of Giessen (1847); studied under Liebig, Dumas, Gerhardt, and Williamson; assistant to Stenhouse, London; professor, Heidelberg (1855), Ghent (1858), and Bonn (1867).

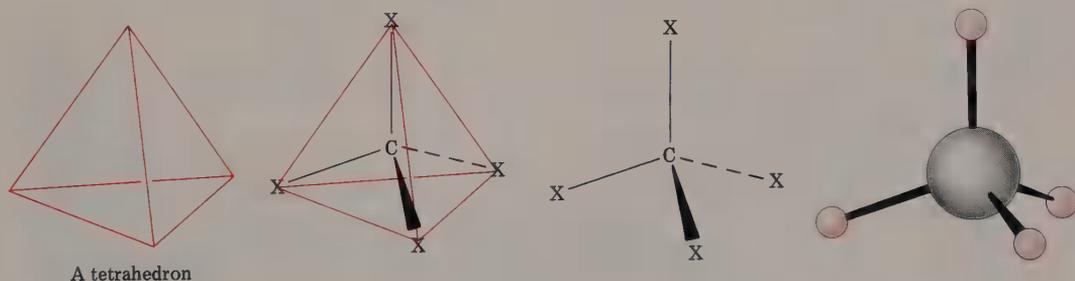
<sup>10</sup>Archibald Scott Couper (1831–1892); b. Kirkintilloch, Scotland; studied at the universities of Glasgow and Edinburgh (1852) and with Würtz in Paris; assistant in Edinburgh (1858).

That is, carbon is *tetravalent*; it always forms four bonds when it joins other elements to form stable compounds. Furthermore, said Kekulé, carbon atoms can bond to each other to form extended chains of carbon atoms linked together.

Shortly after the tetravalent nature of carbon was proposed, extensions to the Kekulé–Couper theory were made when the possibility of *multiple* bonding between atoms was suggested. Emil Erlenmeyer<sup>11</sup> proposed a carbon-to-carbon triple bond for acetylene, and Alexander Crum Brown<sup>12</sup> proposed a carbon-to-carbon double bond for ethylene. In 1865, Kekulé provided another major advance in bonding theory when he postulated that carbon chains can double back on themselves to form rings of atoms.

Perhaps the most significant early advance in the bonding theory of organic molecules was the contribution made independently by Jacobus van't Hoff<sup>13</sup> and Joseph Le Bel.<sup>14</sup> Although Kekulé was correct in describing the tetravalent nature of carbon, chemistry was still viewed in a two-dimensional way until 1874. In that year, van't Hoff and Le Bel added a third dimension to our ideas about molecules by proposing that the four bonds of carbon are not oriented randomly but have a specific spatial direction. Van't Hoff went even further and suggested that the four atoms to which carbon is bonded sit at the corners of a tetrahedron, with carbon in the center.

A representation of a tetrahedral carbon atom is shown in Figure 1.6. Note the conventions used to show three-dimensionality: Solid lines represent bonds in the plane of the page, heavy wedged lines represent bonds



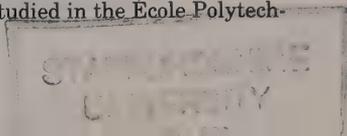
**Figure 1.6** Van't Hoff's tetrahedral carbon atom. The heavy wedged line comes out of the plane of the paper; the normal lines are in the plane; and the dashed line goes back behind the plane of the page.

<sup>11</sup>Richard A. C. E. Erlenmeyer (1825–1909); b. Wehen, Germany; studied in Giessen and in Heidelberg; professor, Munich Polytechnicum (1868–1883).

<sup>12</sup>Alexander Crum Brown (1838–1922); b. Edinburgh; studied at Edinburgh, Heidelberg, and Marburg; professor, Edinburgh (1869–1908).

<sup>13</sup>Jacobus Hendricus van't Hoff (1852–1911); b. Rotterdam; studied at Polytechnic at Delft, Leyden, Bonn, Paris, and received doctorate at Utrecht (1874); professor, Utrecht, Amsterdam (1878–1896), Berlin; Nobel Prize (1901).

<sup>14</sup>Joseph Achille Le Bel (1847–1930); b. Pêchebronn, Alsace; studied in the École Polytechnique and at the Sorbonne; industrial consultant.



coming out of the page toward the viewer, and dashed lines represent bonds receding back behind the page, away from the viewer. These representations will be used throughout the text.

PROBLEM .....

- 1.3 Draw a molecule of chloroform,  $\text{CHCl}_3$ , using wedged, normal, and dashed lines to show its tetrahedral geometry.
- .....

## 1.5 The Nature of Chemical Bonding: Ionic Bonds

---

Why do atoms bond together, and how does our quantum mechanical view of the atom describe bonding? The *why* question is relatively easy to answer: Atoms bond together because the compound that results is more stable (has less energy) than the separate atoms themselves. Just as water always flows downhill, energy is always *released* and flows out of the system when a chemical bond is formed. Conversely, energy is *absorbed* and flows into the system when a chemical bond is broken. The *how* question is more difficult. To answer it, we need to know more about the properties of atoms.

We know through observation that eight electrons (an octet) in the outermost shell (the **valence shell**) impart special stability to the noble-gas elements in Group 8A: Ne (2 + 8); Ar (2 + 8 + 8); Kr (2 + 8 + 18 + 8). We also know that the chemistry of many main-group elements is governed by a tendency to take on the stable electronic makeup of a noble gas. The alkali metals in Group 1A, for example, have single *s* electrons in their valence shells. By losing this electron, they can achieve a noble-gas configuration.

The amount of energy it takes to pull an electron away from an isolated atom in the gas phase is called the **ionization energy** of the element, abbreviated  $E_i$ . Alkali metals, at the far left of the periodic table, give up a valence electron easily, have low ionization energies, and are said to be **electropositive**. Elements at the middle and far right of the periodic table hold their valence electrons more tightly, give them up less readily, and therefore have higher  $E_i$ 's. In other words, a low  $E_i$  corresponds to the ready loss of an electron, and a high  $E_i$  corresponds to the difficult loss of an electron. Table 1.3 lists the ionization energies of the second-row elements in kilocalories per mole (kcal/mol), and Figure 1.7 displays  $E_i$  data for the entire periodic table.

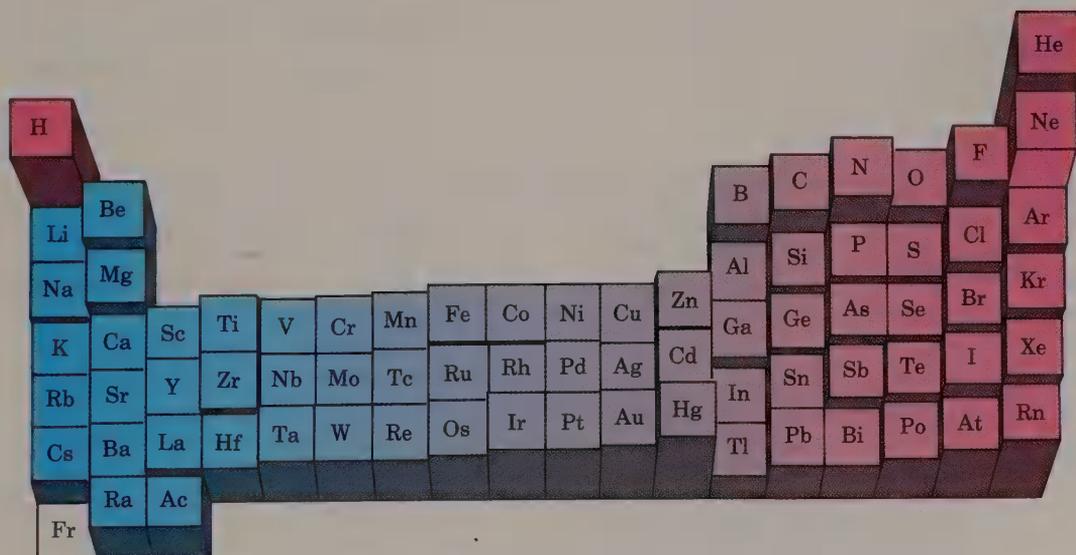


Just as the electropositive alkali metals at the left of the periodic table have a tendency to form positive ions by losing an electron, the halogens (Group 7A elements) at the right of the periodic table have a tendency to form *negative* ions by *gaining* an electron. By so doing, the halogens can

TABLE 1.3 Ionization Energies of Some Elements<sup>a</sup>

<i>Element (electronic configuration)</i>		<i>Cation (electronic configuration)</i>	<i>Ionization energy (kcal/mol)</i>
Li (1s <sup>2</sup> 2s)	→ <sup>-e<sup>-</sup></sup>	Li <sup>+</sup> (1s <sup>2</sup> —same as He)	125
Na (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s)	→ <sup>-e<sup>-</sup></sup>	Na <sup>+</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> —same as Ne)	118
K (... 3s <sup>2</sup> 3p <sup>6</sup> 4s)	→ <sup>-e<sup>-</sup></sup>	K <sup>+</sup> (... 3s <sup>2</sup> 3p <sup>6</sup> —same as Ar)	100
C (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>2</sup> )	→ <sup>-e<sup>-</sup></sup>	C <sup>+</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p)	259
C <sup>+</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p)	→ <sup>-e<sup>-</sup></sup>	C <sup>2+</sup> (1s <sup>2</sup> 2s <sup>2</sup> )	562
F (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup> )	→ <sup>-e<sup>-</sup></sup>	F <sup>+</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>4</sup> )	401
Ne (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> )	→ <sup>-e<sup>-</sup></sup>	Ne <sup>+</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup> )	497

<sup>a</sup>Organic chemists have been slow to adopt SI units, preferring to use kilocalories (kcal), rather than kilojoules (kJ), as a measure of energy. This book will use dual units, with values shown in both kcal/mol and kJ/mol, where 1 calorie = 4.184 joule, or 1 kcal/mol = 4.184 kJ/mol.



**Figure 1.7** A graphic display of relative ionization energies. Metallic elements on the left of the table have lower ionization energies than do nonmetallic elements on the right.

achieve noble-gas configurations. The measure of this tendency of an isolated atom in the gas phase to gain an electron is called the atom's **electron affinity**, abbreviated  $E_{ea}$ . Energy is released when an electron is added to most elements, and  $E_{ea}$ 's therefore have negative values.

Nonmetals on the right side of the periodic table have a much greater tendency to add an electron than do metals on the left side and are said to be **electronegative**. Thus, the halogens release a large amount of energy when they add an electron and have much larger negative electron affinities than do the alkali metals. Table 1.4 lists the electron affinities of some common elements.



**TABLE 1.4** Electron Affinities of Some Elements

<i>Element (electronic configuration)</i>		<i>Anion (electronic configuration)</i>	<i>Electron affinity (kcal/mol)</i>
Li (1s <sup>2</sup> 2s)	+e <sup>-</sup> →	Li <sup>-</sup> (1s <sup>2</sup> 2s <sup>2</sup> )	-13.6
Na (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s)	+e <sup>-</sup> →	Na <sup>-</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> 3s <sup>2</sup> )	-5.0
C (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>2</sup> )	+e <sup>-</sup> →	C <sup>-</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>3</sup> )	-28.9
F (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>5</sup> )	+e <sup>-</sup> →	F <sup>-</sup> (1s <sup>2</sup> 2s <sup>2</sup> 2p <sup>6</sup> —same as Ne)	-79.6
Cl (... 3s <sup>2</sup> 3p <sup>5</sup> )	+e <sup>-</sup> →	Cl <sup>-</sup> (... 3s <sup>2</sup> 3p <sup>6</sup> —same as Ar)	-83.2

The simplest kind of chemical bonding is that between an electropositive element (low  $E_i$ ) and an electronegative element (large negative  $E_{ea}$ ). For example, when sodium metal [ $E_i = 118.5$  kcal/mol (496 kJ/mol)] reacts with chlorine gas [ $E_{ea} = -83.3$  kcal/mol (-349 kJ/mol)], sodium can donate an electron to chlorine forming positively charged sodium ions and negatively charged chloride ions. When a vast number of sodium atoms transfer electrons to an equal number of chlorine atoms, a visible crystal of sodium chloride results. The NaCl product is said to have **ionic bonding** and to be an **ionic solid**. That is, the ions are held together to a large extent by electrostatic attraction between the two unlike charges. A similar situation exists for many other metal salts such as potassium fluoride (K<sup>+</sup>F<sup>-</sup>), lithium bromide (Li<sup>+</sup>Br<sup>-</sup>), and so on. This picture of the ionic bond, first proposed by Walter Kössel<sup>15</sup> in 1916, satisfactorily accounts for the chemistry of many inorganic compounds.

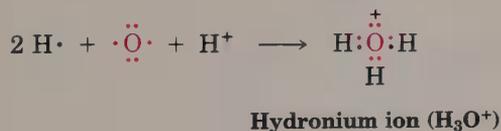
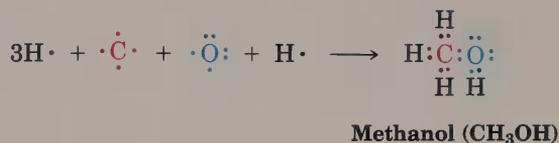
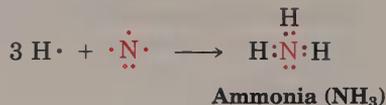
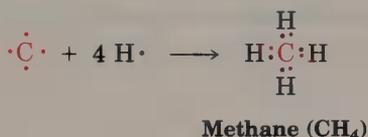
<sup>15</sup>Walter Ludwig Julius Paschen Heinrich Kössel (1888–1956); b. Berlin; assistant in physics in Heidelberg (1910) and Munich (1913); professor of physics, Kiel (1921), Danzig (Poland) (1932–1945), and Tübingen (1947).

## 1.6 The Nature of Chemical Bonding: Covalent Bonds

We've just seen that elements on the far left (sodium) and far right (chlorine) of the periodic table form ionic bonds by gaining or losing an electron to achieve a noble-gas configuration. How, though, do elements in the middle of the periodic table form bonds? Let's look at the carbon atom in methane,  $\text{CH}_4$ , as an example. Certainly the bonding in methane isn't ionic, because it would be very difficult for carbon ( $1s^2 2s^2 2p^2$ ) either to gain or to lose *four* electrons to achieve a noble-gas configuration.<sup>16</sup>

In fact, carbon bonds to other atoms, not by donating electrons but by *sharing* them. Such shared-electron bonds, first proposed in 1916 by G. N. Lewis,<sup>17</sup> are called **covalent bonds**, and the collection of atoms held together by covalent bonds is called a **molecule**. The covalent bond is the most important bond in organic chemistry.

A simple shorthand way of indicating covalent bonds in molecules is to use **Lewis structures**, or **electron-dot structures**, in which the valence electrons of an atom are represented as dots. Thus, hydrogen has one dot representing its  $1s$  electron, carbon has four dots ( $2s^2 2p^2$ ), oxygen has six dots ( $2s^2 2p^4$ ), and so on. A stable molecule results whenever a noble-gas configuration with filled  $s$  and  $p$  valence orbitals is achieved for all atoms, as in the following examples:

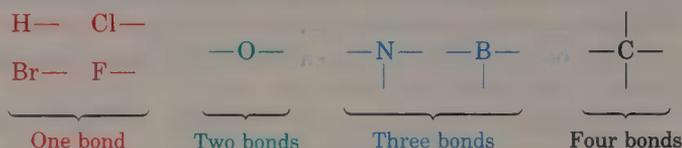


The number of covalent bonds an atom forms depends on how many valence electrons it has. Atoms with one, two, or three valence electrons form one, two, or three bonds, but atoms with four or more valence-shell

<sup>16</sup>The electronic configuration of carbon can be written either as  $1s^2 2s^2 2p^2$  or as  $1s^2 2s^2 2p_x 2p_y$ . Both notations are correct, but the latter is more informative because it indicates that two of the three equivalent  $p$  orbitals are half filled.

<sup>17</sup>Gilbert Newton Lewis (1875–1946); b. Weymouth, Mass.; Ph.D. Harvard (1899); professor, Massachusetts Institute of Technology (1905–1912), University of California, Berkeley (1912–1946).

electrons form as many bonds as they need electrons to fill the  $s$  and  $p$  levels of their valence shells. Thus, boron has three valence electrons ( $2s^2 2p^1$ ) and forms three bonds, as in  $\text{BH}_3$ ; carbon ( $2s^2 2p^2$ ) fills its valence shell and achieves the configuration of neon by forming four bonds, as in  $\text{CH}_4$ ; nitrogen ( $2s^2 2p^3$ ) achieves the neon configuration by forming three bonds, as in  $\text{NH}_3$ ; and oxygen ( $2s^2 2p^4$ ) achieves the neon configuration by forming two bonds, as in  $\text{H}_2\text{O}$ .



Lewis structures are valuable because they make electron bookkeeping possible and act as reminders of the number of valence electrons present. Simpler still is the use of “**Kekulé**” structures, also called **line-bond structures**, in which a two-electron covalent bond is indicated simply as a line drawn between atoms. Pairs of nonbonding valence electrons are often ignored when drawing line-bond structures, but it’s still necessary to keep mentally aware of their existence, and it’s useful when starting out to always include them. Some examples of these structures are shown in Table 1.5.

**TABLE 1.5** Lewis and Kekulé Structures of Some Simple Molecules

Name	Lewis structure	Kekulé structure	Name	Lewis structure	Kekulé structure
Water ( $\text{H}_2\text{O}$ )	$\text{H}:\ddot{\text{O}}:\text{H}$	$\text{H}-\text{O}-\text{H}$	Methane ( $\text{CH}_4$ )	$\begin{array}{c} \text{H} \\ \vdots \\ \text{H}:\text{C}:\text{H} \\ \vdots \\ \text{H} \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{H}-\text{C}-\text{H} \\   \\ \text{H} \end{array}$
Ammonia ( $\text{NH}_3$ )	$\begin{array}{c} \text{H} \\ \vdots \\ \text{H}:\text{N}:\text{H} \\ \vdots \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{H}-\text{N}-\text{H} \\   \\ \text{H} \end{array}$	Methanol ( $\text{CH}_3\text{OH}$ )	$\begin{array}{c} \text{H} \\ \vdots \\ \text{H}:\text{C}:\ddot{\text{O}}:\text{H} \\ \vdots \\ \text{H} \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{H}-\text{C}-\text{O}-\text{H} \\   \\ \text{H} \end{array}$

**PRACTICE PROBLEM**.....

How many hydrogen atoms does phosphorus bond to in forming phosphine,  $\text{PH}_x$ ?

**Solution** Because phosphorus is in Group 5A of the periodic table, it has five valence electrons. It needs to share three more electrons to make an octet, and therefore bonds to three hydrogen atoms, giving  $\text{PH}_3$ .

**PROBLEM**.....

1.4 What are likely formulas for these substances?

- (a)  $\text{CCl}_x$     (b)  $\text{AlH}_x$     (c)  $\text{CH}_2\text{Cl}_2$     (d)  $\text{SiF}_x$     (e)  $\text{CH}_3\text{NH}_x$

## PROBLEM.....

- 1.5 Write both Lewis and line-bond structures for these substances, showing all non-bonded electrons:
- |  |   |
|--|---|
| (a) $\text{CHCl}_3$ , chloroform           | (b) $\text{H}_2\text{S}$ , hydrogen sulfide |
| (c) $\text{CH}_3\text{NH}_2$ , methylamine | (d) $\text{BH}_3$ , borane                  |
| (e) $\text{NaH}$ , sodium hydride          | (f) $\text{CH}_3\text{Li}$ , methyllithium  |

## PROBLEM.....

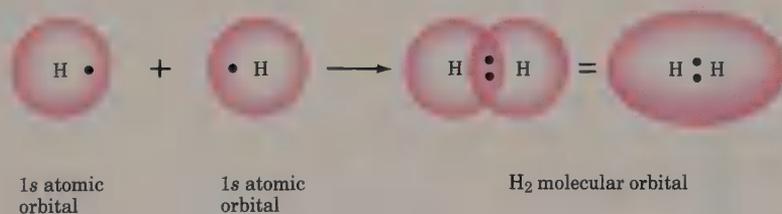
- 1.6 Which of these substances would you expect to have covalent bonds and which ionic bonds? Explain.
- (a)  $\text{CH}_4$     (b)  $\text{CH}_2\text{Cl}_2$     (c)  $\text{LiI}$     (d)  $\text{KBr}$     (e)  $\text{MgCl}_2$     (f)  $\text{Cl}_2$

## PROBLEM.....

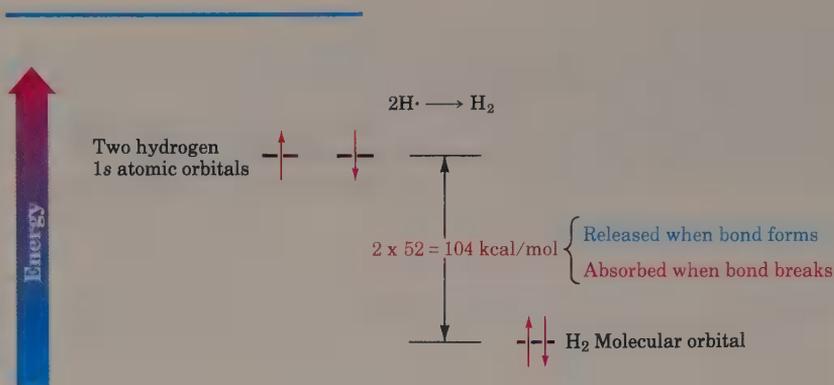
- 1.7 Write both a Lewis structure and a line-bond structure for ethane,  $\text{C}_2\text{H}_6$ .
- .....

## 1.7 Formation of Covalent Bonds: Molecular Orbital Theory

How are covalent bonds formed? The most generally satisfactory method for dealing with covalent bonds in organic compounds is **molecular orbital (MO) theory**. Molecular orbital theory pictures the formation of covalent bonds by an *overlapping* of atomic orbitals. For example, we can describe the hydrogen molecule ( $\text{H}-\text{H}$ ) by imagining what might happen if two hydrogen atoms, each with an atomic  $1s$  orbital, meet and join together. As the two spherical *atomic* orbitals approach each other and combine, a new egg-shaped *molecular* orbital results, so called because it belongs to the entire  $\text{H}_2$  molecule rather than to an individual atom. The molecular orbital is filled by two electrons, one donated by each hydrogen:

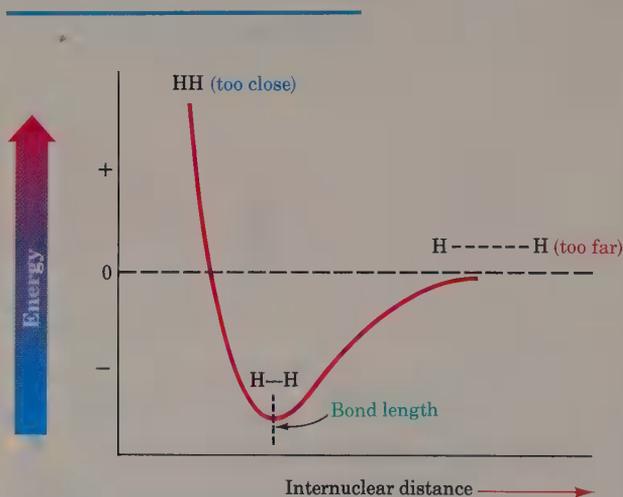


During the reaction  $2 \text{H} \cdot \rightarrow \text{H}_2$ , 104 kcal/mol (435 kJ/mol) of energy is released. Since the product  $\text{H}_2$  molecule has 104 kcal/mol less energy than the starting  $2 \text{H} \cdot$ , we say that the product is more stable than the starting material and that the new  $\text{H}-\text{H}$  bond has a **bond strength** of 104 kcal/mol. In other words, we would have to put 104 kcal/mol of energy *into* the  $\text{H}-\text{H}$  bond in order to break the hydrogen molecule apart into two hydrogen atoms. Figure 1.8 shows the relative energy levels of the different orbitals.



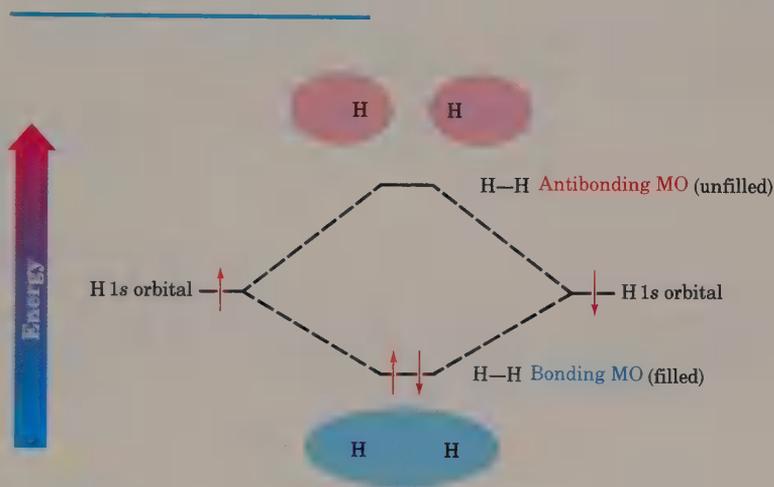
**Figure 1.8** Energy levels of H atomic orbitals and  $\text{H}_2$  molecular orbital. The  $\text{H}_2$  molecular orbital is lower in energy by 52 kcal/mol ( $2 \times 52 = 104$ ).

How close are the two nuclei in the hydrogen molecule? If they're *too* close together, they'll repel each other electrostatically since both are positively charged. Yet if the nuclei are too far apart, they won't be able to share the bonding electrons adequately. Thus, there is an optimum distance between the two nuclei that leads to maximum stability (Figure 1.9). This optimum distance is called the **bond length** and has a value of 0.74 Å in the hydrogen molecule. Every covalent bond has both a characteristic bond strength and bond length.



**Figure 1.9** A plot of energy versus internuclear distance for two hydrogen atoms. The distance at the lowest energy point is called the bond length.

One further point about this description of the hydrogen molecule is that an orbital seems to have disappeared. We began forming the hydrogen molecule by combining two atomic orbitals, each of which, if filled, could have held two electrons, for a total of four. We ended up, however, with what seems to be *one* molecular orbital, which can hold only two electrons. In fact, an orbital hasn't disappeared; we simply haven't paid it much attention. When we combine a pair of atomic orbitals, a pair of molecular orbitals is produced. One of the molecular orbitals is lower in energy than the starting atomic orbitals, and the other is correspondingly higher in energy, as shown in Figure 1.10.

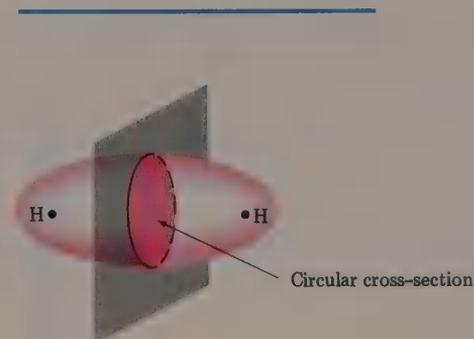


**Figure 1.10** Molecular orbitals of H<sub>2</sub>. The combination of two hydrogen 1s atomic orbitals leads to the formation of two molecular orbitals. The lower-energy (bonding) molecular orbital is filled, and the higher-energy (antibonding) molecular orbital is unfilled.

The lower-energy orbital is called a **bonding MO**, and the higher-energy orbital is called an **antibonding MO**. The two electrons occupy the low-energy bonding orbital, with the high-energy antibonding orbital left unfilled. After all, if electrons were to occupy this high-energy orbital, the molecule would be higher in energy than the two isolated hydrogen atoms, and no bond could result—thus the term *antibonding*.

The bonding molecular orbital in the hydrogen molecule has the elongated egg shape that we might get by pressing two spheres together. If an imaginary plane were to pass through the middle of the orbital, the intersection of the plane and the orbital would look like a circle. In other words, the H–H bond is cylindrically symmetrical, as shown in Figure 1.11.

Bonds that have circular cross-sections and are formed by head-on overlap of two atomic orbitals are called **sigma ( $\sigma$ ) bonds**. Although sigma



**Figure 1.11** The cylindrical symmetry of the H–H bond. The intersection of a plane cutting through the orbital looks like a circle.

bonds are the most common kind, there are other types of bonds as well. Let's consider the fluorine molecule,  $F_2$ , for example. A fluorine atom has seven valence electrons and the electronic configuration  $1s^2 2s^2 2p^5$ . By bonding together, two fluorine atoms can each achieve stable valence octets:

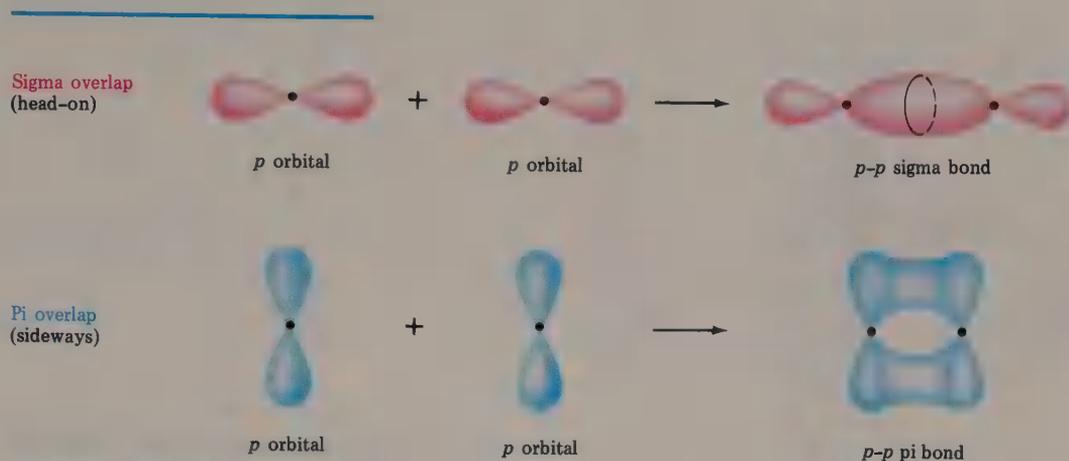


Unlike the situation in the hydrogen atom, though, a fluorine atom has an unshared  $2p$  electron rather than a  $1s$  electron. How can two  $p$  orbitals come together to form a bond?

The general answer to this question was provided by Linus Pauling<sup>18</sup> in 1931 when he stated the **principle of maximum orbital overlap**. According to this principle, the strongest bond is formed when the two orbitals achieve maximum overlap. There are two geometric possibilities for  $p$ -orbital overlap in the fluorine molecule: The  $p$  orbitals can overlap in a head-on fashion to form a sigma bond, or they can overlap in a sideways manner to form what is called a **pi ( $\pi$ ) bond**, shown in Figure 1.12.

Though it's difficult to predict which kind of bonding leads to maximum overlap, it turns out that sigma bonding is usually more efficient than pi bonding. Fluorine therefore forms a sigma molecular bond between two  $2p$  orbitals. The new F–F bond has a bond strength of 38 kcal/mol (159 kJ/mol) and a bond length of 1.42 Å.

<sup>18</sup>Linus Pauling (1901– ); b. Portland, Oregon; Ph.D. California Institute of Technology (1925); professor, California Institute of Technology (1925–1967); University of California, San Diego; Professor Emeritus, Stanford University (1974– ); Nobel Prize (1954, 1963).



**Figure 1.12** The formation of sigma and pi bonds by overlap of  $p$  orbitals.

**PROBLEM** .....

- 1.8 We said that sigma bonds, formed by head-on overlap of orbitals, have cylindrical symmetry. Draw a cross-section of a pi bond, formed by sideways overlap of  $p$  orbitals. How does pi bond symmetry differ from sigma bond symmetry?
- .....

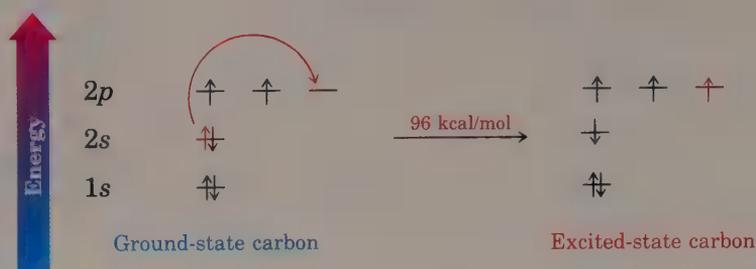
## 1.8 Hybridization: $sp^3$ Orbitals and the Structure of Methane

The bonding in both the hydrogen molecule and the fluorine molecule is fairly straightforward, but the situation becomes more complicated when we turn to organic molecules with tetravalent carbon atoms. Let's start with the simplest case and consider methane,  $\text{CH}_4$ . Carbon has four electrons in its valence shell, two of which are paired in the  $2s$  orbital and two of which are unpaired in different  $2p$  orbitals. Thus, the ground-state electronic configuration of carbon is  $1s^2 2s^2 2p_x 2p_y$ .

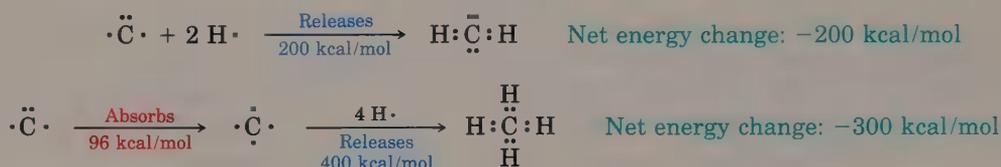
The first problem is immediately apparent: Carbon needs to share four electrons to make four bonds, but only two of its valence electrons are unpaired. The two  $2s$  electrons are already paired with each other in the same orbital and can't be shared in bonds with other atoms. Why, therefore, doesn't carbon bond to two hydrogen atoms to form  $\text{CH}_2$ ? In fact,  $\text{CH}_2$  is a known compound, but it is highly reactive and has only a fleeting existence.

To get around the difficulty in having enough unpaired electrons, carbon must adopt an electronic configuration different from the ground-state configuration. By promoting one electron from the filled  $2s$  orbital into the

vacant  $2p_z$  orbital, carbon can achieve the higher-energy configuration  $1s^2 2s^2 2p_x 2p_y 2p_z$ .



Carbon has four unpaired electrons in its higher-energy configuration and can now form four bonds to hydrogens. Although 96 kcal/mol (402 kJ/mol) is required to promote the 2s electron to a 2p orbital, this energy loss is more than offset by the formation of four stable C–H bonds rather than two. Since the formation of each C–H bond releases about 100 kcal/mol of energy, a net release of nearly 300 kcal/mol occurs when  $\text{CH}_4$  is formed, versus a release of only 200 kcal/mol when  $\text{CH}_2$  is formed.

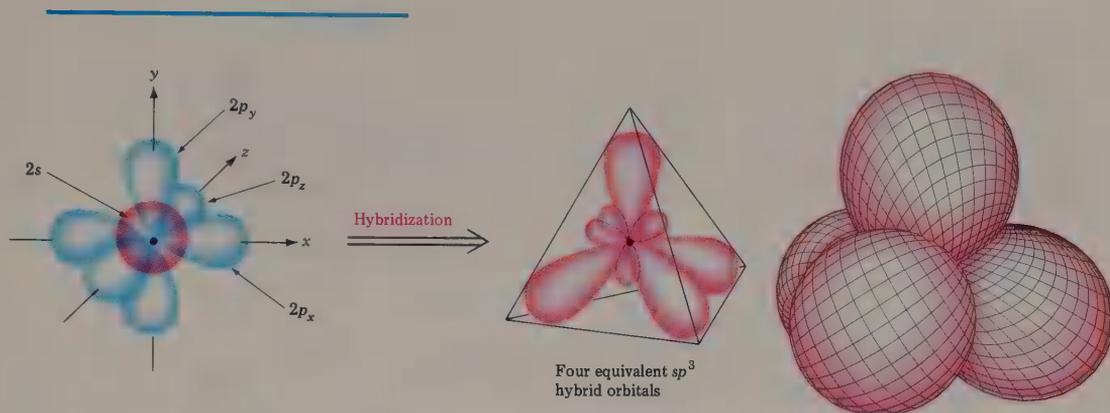


What are the four C–H bonds in methane like? Since carbon uses two kinds of orbitals (2s and 2p) for bonding purposes, we might expect methane to have two kinds of C–H bonds. In fact, though, a large amount of evidence shows that all four C–H bonds in methane are identical. How can we explain this?

The answer was provided in 1931 by Linus Pauling, who showed that an s orbital and three p orbitals can mathematically combine, or **hybridize**, to form four equivalent atomic orbitals that are spatially oriented toward the corners of a tetrahedron. Shown in Figure 1.13, these new tetrahedral orbitals are called  **$sp^3$  hybrids**<sup>19</sup> because they arise mathematically from a combination of three p orbitals and one s orbital.

The concept of hybridization explains *how* carbon forms four equivalent tetrahedral bonds but doesn't answer the question of why it does so. Viewing a cross-section of an  $sp^3$  hybrid orbital suggests the answer. When an s orbital hybridizes with three p orbitals, the resultant hybrid orbitals are unsymmetrical about the nucleus. One of the two lobes of an  $sp^3$  orbital is much larger than the other (Figure 1.14) and can therefore overlap better with another orbital when it forms a bond. As a result,  $sp^3$  hybrid orbitals form stronger bonds than do unhybridized s or p orbitals.

<sup>19</sup>Note that the superscript used to identify an  $sp^3$  hybrid orbital tells how many of each type of atomic orbital combine to form the hybrid, not how many electrons occupy that orbital.



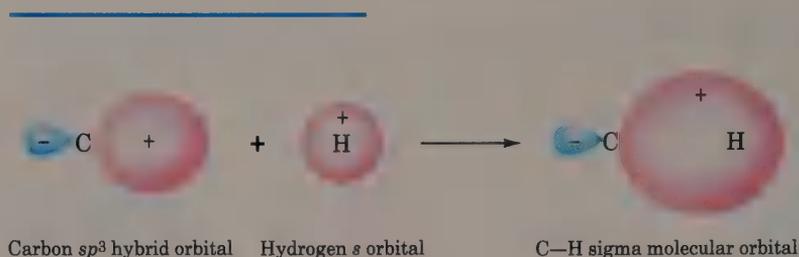
**Figure 1.13** Four  $sp^3$  hybrid orbitals, formed by the combination of an atomic  $s$  orbital and three atomic  $p$  orbitals.



**Figure 1.14** The formation of an  $sp^3$  hybrid orbital by overlap of a  $p$  orbital with part of an  $s$  orbital. Overlap of the  $s$  orbital with the positive  $p$  lobe is additive, but overlap with the negative  $p$  lobe cancels out. The resultant hybrid orbital is strongly oriented in one direction.

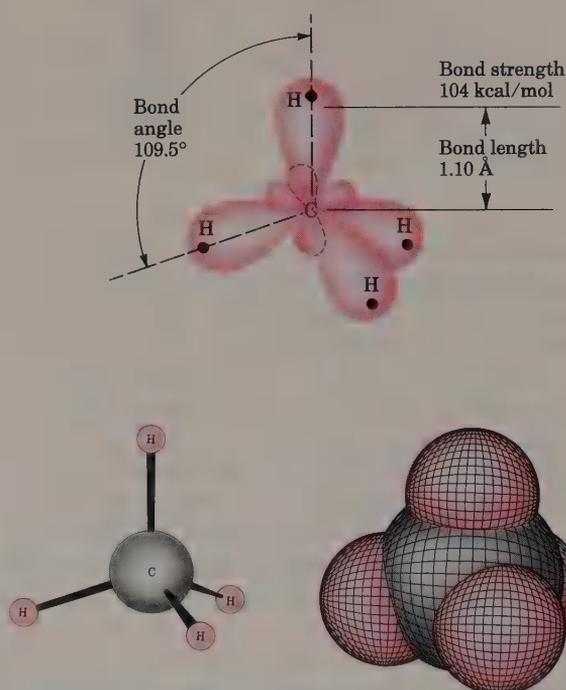
This lack of symmetry in  $sp^3$  orbitals arises because of a property of orbitals that we've not yet considered. When the wave equation for a  $p$  orbital is solved, the two lobes have opposite algebraic signs, + and -. Thus, when a  $p$  orbital hybridizes with an  $s$  orbital, one lobe is *additive* with the  $s$  orbital, but the other lobe is *subtractive*. The resultant hybrid orbital is therefore strongly oriented in one direction.

We describe the  $sp^3$  hybrid as a *directed* orbital, and we find that it is capable of forming strong bonds by overlapping the orbitals of other atoms. For example, the overlap of a carbon  $sp^3$  hybrid orbital with a hydrogen  $1s$  orbital gives a strong C-H bond (Figure 1.15).



**Figure 1.15** The formation of a C–H bond by head-on ( $\sigma$ ) overlap of a carbon  $sp^3$  hybrid orbital with a hydrogen  $1s$  orbital.

When the four identical orbitals of an  $sp^3$  hybridized carbon atom overlap with four hydrogen atoms, four identical C–H bonds are formed. Each C–H bond in methane has a strength of 104 kcal/mol (435 kJ/mol) and a length of 1.10 Å. Since the four bonds have a specific geometry, we can also define a third important property called the **bond angle**. The angle formed by each H–C–H is exactly  $109.5^\circ$ , the so-called tetrahedral angle. Methane therefore has the structure shown in Figure 1.16.



**Figure 1.16** The structure of methane. The second and third drawings are computer generated.

Before ending this discussion of the methane structure, it should again be pointed out that no orbitals are “lost” during hybridization and C–H bond formation. An  $sp^3$ -hybridized carbon has four hybrid orbitals and four electrons. Four hydrogen atoms also provide four  $1s$  orbitals and four electrons. Methane therefore has four bonding C–H molecular orbitals, which are filled, and four antibonding C–H orbitals, which are unfilled. All eight MOs are shown in Figure 1.17.

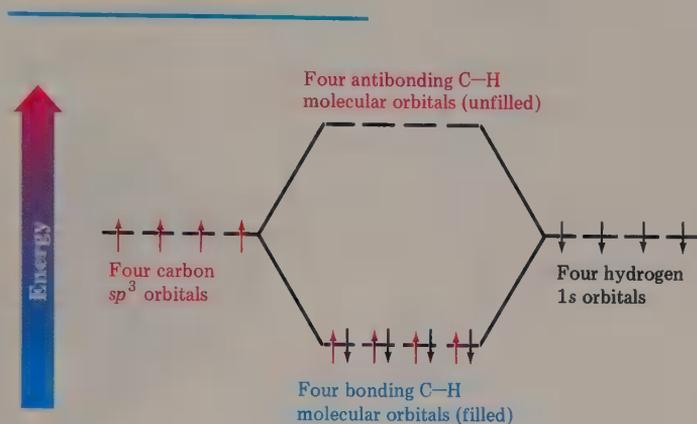
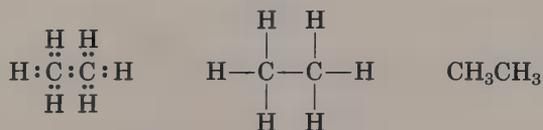


Figure 1.17 Molecular orbitals of methane.

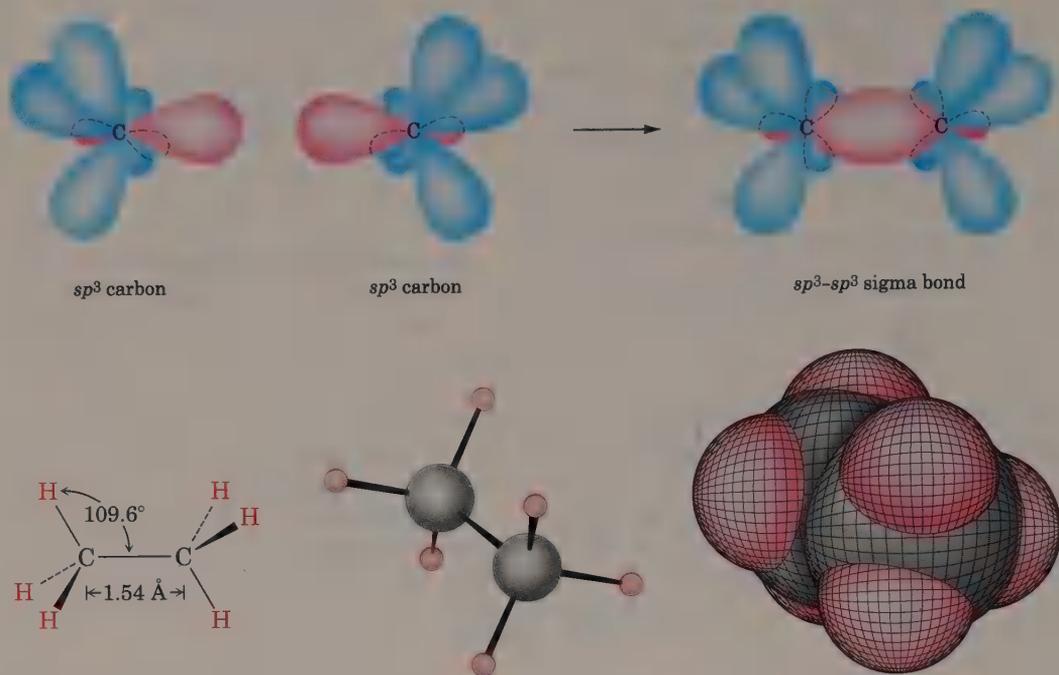
## 1.9 The Structure of Ethane

A special characteristic of carbon is that it can form stable bonds to other carbon atoms. The same kind of hybridization that explains the methane structure also explains how one carbon atom can bond to another to make possible the many millions of known organic compounds. Ethane,  $C_2H_6$ , is the simplest molecule containing a carbon–carbon bond:



Some representations of ethane

We can picture the ethane molecule by imagining that the two carbon atoms bond to each other by sigma overlap of an  $sp^3$  hybrid orbital from each. The remaining three  $sp^3$  hybrid orbitals on each carbon overlap with hydrogen orbitals to form the six C–H bonds, as shown in Figure 1.18. The C–H bonds in ethane are similar to those in methane, though a bit weaker (98 kcal/mol for ethane versus 104 kcal/mol for methane). The C–C bond



**Figure 1.18** The structure of ethane. The carbon–carbon bond is formed by sigma overlap of two carbon  $sp^3$  hybrid orbitals.

is 1.54 Å long and has a strength of 88 kcal/mol (368 kJ/mol). All the bond angles of ethane are very near the tetrahedral value, 109.5°.

**PROBLEM** .....

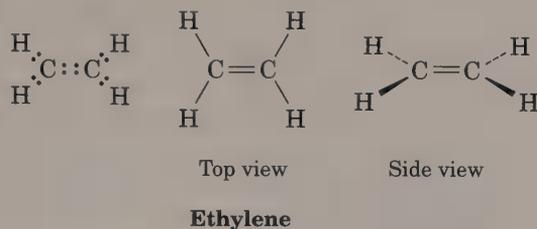
- 1.9 Draw a line-bond structure for propane,  $\text{CH}_3\text{CH}_2\text{CH}_3$ . Predict the value of each bond angle, and indicate the overall shape of the molecule.

**PROBLEM** .....

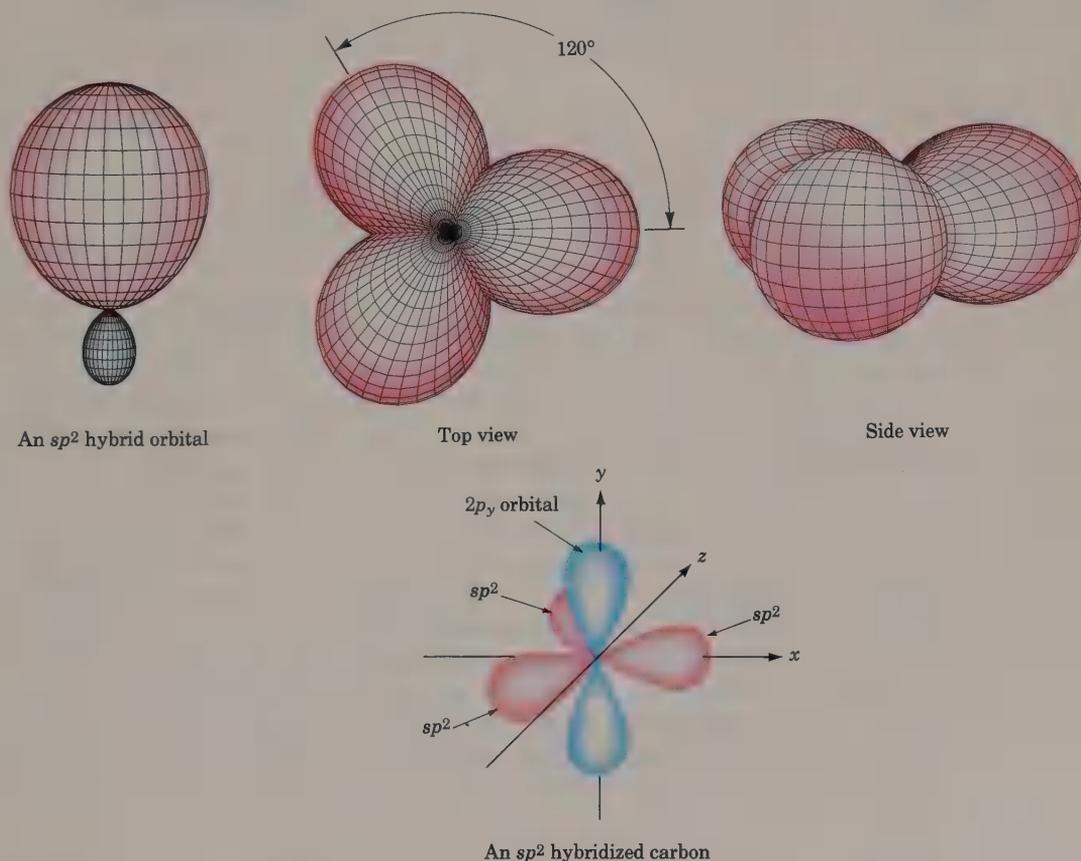
- 1.10 Why can't an organic molecule have the formula  $\text{C}_2\text{H}_7$ ?  
.....

## 1.10 Hybridization: $sp^2$ Orbitals and the Structure of Ethylene

Although  $sp^3$  hybridization is the most common electronic state of carbon, it's not the only possibility. Let's look at ethylene,  $\text{C}_2\text{H}_4$ , for example. It was recognized over 100 years ago that ethylene carbons can be tetravalent only if the two carbon atoms are linked by a *double bond*.



When we formed  $sp^3$  hybrid orbitals to explain the bonding in methane, we first promoted an electron from the  $2s$  orbital of ground-state carbon to form a higher-energy configuration with four unpaired electrons. We then mathematically combined the four singly occupied atomic orbitals to construct four  $sp^3$  hybrids. Imagine instead that we mathematically combine the  $2s$  orbital with only two of the three available  $2p$  orbitals. Three hybrid orbitals called  **$sp^2$  hybrids** result, and one unhybridized  $2p$  orbital remains unchanged. The three  $sp^2$  orbitals lie in a plane at angles of  $120^\circ$  to each other, with the remaining  $p$  orbital perpendicular to the  $sp^2$  plane, as shown in Figure 1.19.



**Figure 1.19** An  $sp^2$ -hybridized carbon. There are three equivalent  $sp^2$  hybrid orbitals lying in a plane at angles of  $120^\circ$  to each other, and a single unhybridized  $p$  orbital perpendicular to the  $sp^2$  plane.

As with  $sp^3$  hybrid orbitals,  $sp^2$  hybrids are strongly oriented in a specific direction and can form strong bonds. If we allow two  $sp^2$ -hybridized carbons to approach each other, they can form a strong sigma bond by  $sp^2$ - $sp^2$  overlap. When this occurs, the unhybridized  $p$  orbitals on each carbon also approach each other with the correct geometry for sideways overlap to form a pi bond. The combination of  $sp^2$ - $sp^2$  sigma overlap and  $2p$ - $2p$  pi overlap results in the net sharing of four electrons and the formation of a carbon-carbon double bond (Figure 1.20).

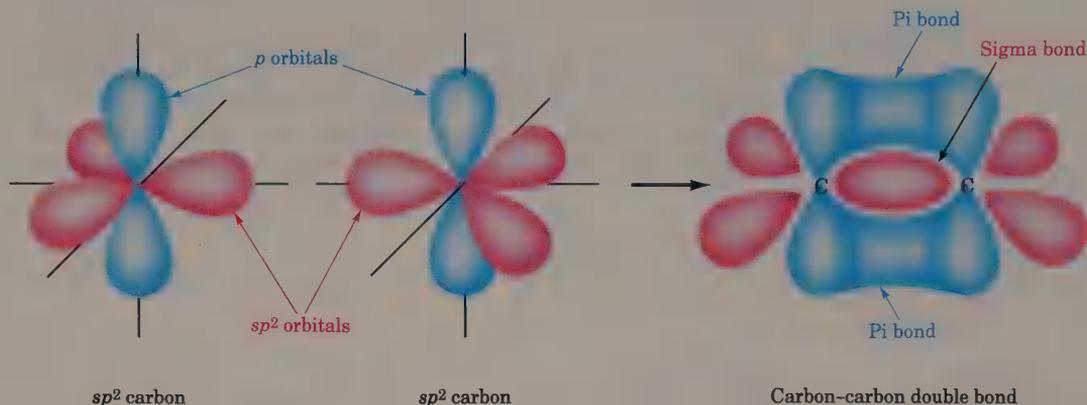


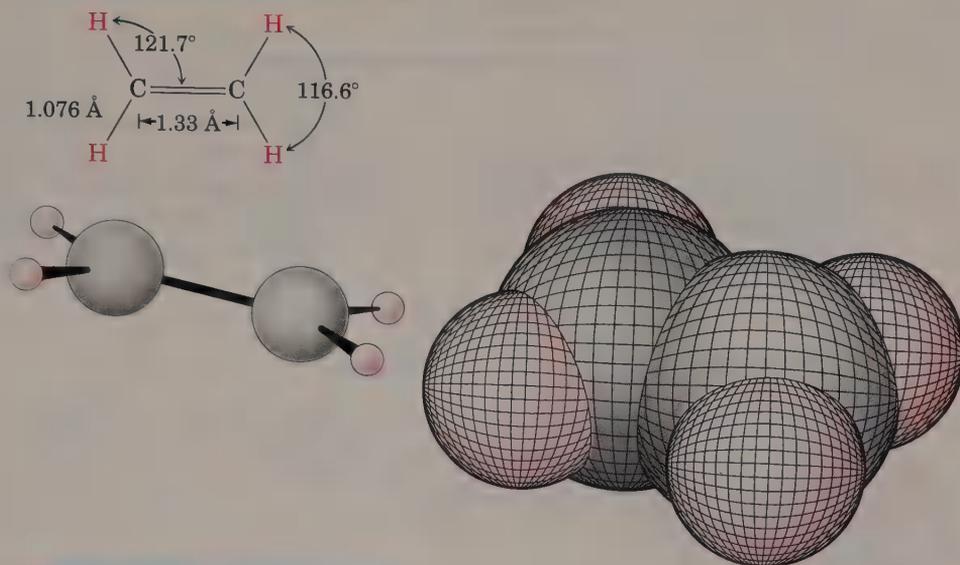
Figure 1.20 Orbital overlap in a carbon-carbon double bond.

To complete the structure of ethylene, we need only allow four hydrogen atoms to sigma bond to the remaining four  $sp^2$  orbitals. The resultant ethylene molecule has a planar (flat) structure with H-C-H and H-C-C bond angles of approximately  $120^\circ$  (the H-C-H bond angles are  $116.6^\circ$ , and the H-C-C bond angles are  $121.7^\circ$ ). Each C-H bond has a length of  $1.076 \text{ \AA}$  and a strength of  $103 \text{ kcal/mol}$  ( $431 \text{ kJ/mol}$ ).

We might expect the central carbon-carbon double bond in ethylene to be both shorter and stronger than the ethane single bond because it results from sharing four electrons rather than two. This expectation has been verified, for ethylene has a C-C bond length of  $1.33 \text{ \AA}$  and a strength of  $152 \text{ kcal/mol}$  ( $636 \text{ kJ/mol}$ ) versus a length of  $1.54 \text{ \AA}$  and a strength of  $88 \text{ kcal/mol}$  for ethane. Note, though, that the strength of a carbon-carbon double bond is not *exactly* twice as large as that of a single bond because overlap in the pi part of the double bond is not as effective as overlap in the sigma part. The structure of ethylene is shown in Figure 1.21.

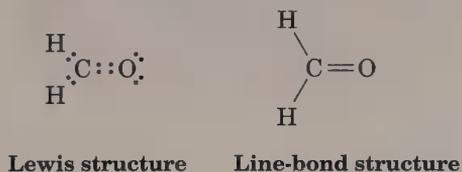
#### PRACTICE PROBLEM.....

Formaldehyde,  $\text{CH}_2\text{O}$ , contains a carbon-oxygen double bond. Draw Lewis and line-bond structures of formaldehyde, and indicate the hybridization of the carbon atom.



**Figure 1.21** The structure of ethylene. Note that the computer-generated structures show only the connections between atoms and do not explicitly show the carbon–carbon double bond.

**Solution** There is only one way that two hydrogens, one carbon, and one oxygen can combine:



Like the carbon atoms in ethylene, the carbon atom in formaldehyde is  $sp^2$  hybridized.

**PROBLEM** .....

- 1.11** Draw all of the bonds in propene,  $\text{CH}_3\text{CH}=\text{CH}_2$ . Indicate the hybridization of each carbon, and predict the value of each bond angle.

**PROBLEM** .....

- 1.12** Answer Problem 1.11 for 1,3-butadiene,  $\text{H}_2\text{C}=\text{CH}-\text{CH}=\text{CH}_2$ .

**PROBLEM** .....

- 1.13** Draw both a Lewis structure and a line-bond structure for acetaldehyde,  $\text{CH}_3\text{CHO}$ .

.....

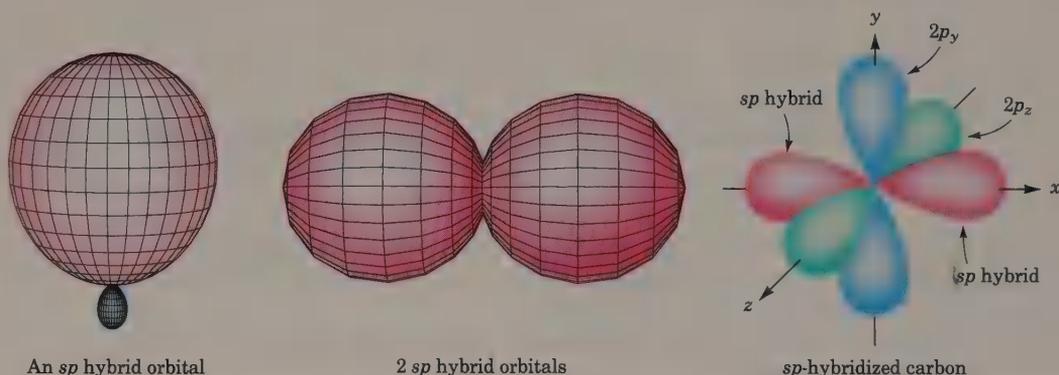
## 1.11 Hybridization: $sp$ Orbitals and the Structure of Acetylene

In addition to being able to form single and double bonds by sharing two and four electrons, carbon can also form a third kind of bond by sharing six electrons. Acetylene,  $C_2H_2$ , can be satisfactorily pictured only if we assume that it contains a carbon-carbon triple bond. To account for the triple bond, we must construct a third kind of hybrid orbital, an  $sp$  hybrid:



Acetylene

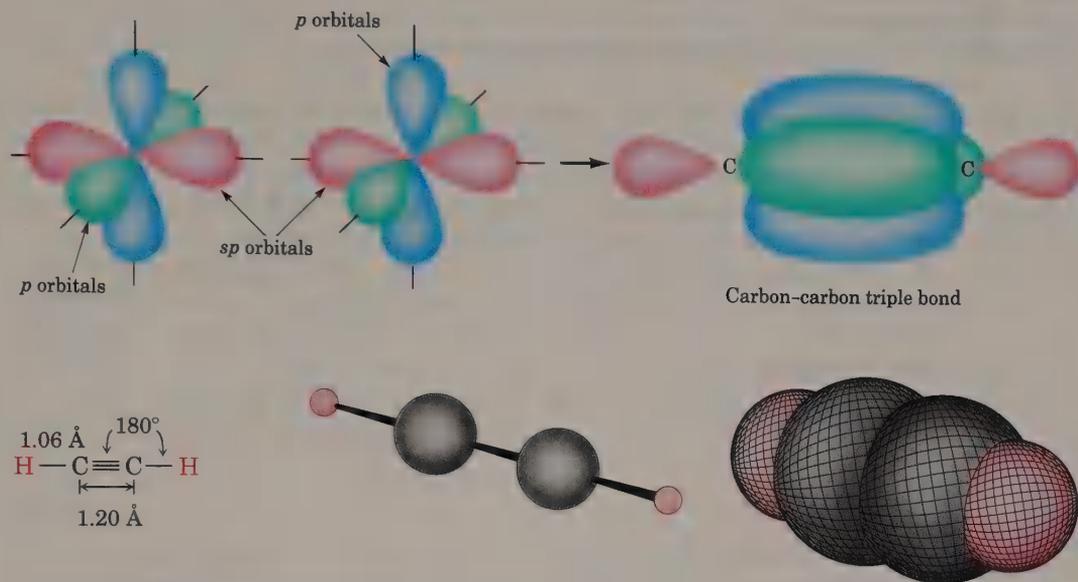
Imagine that, instead of combining with two or three  $p$  orbitals, a carbon  $2s$  orbital hybridizes with only a single  $p$  orbital. Two  $sp$  hybrid orbitals result, and two  $p$  orbitals remain unchanged. The two  $sp$  orbitals are linear, or  $180^\circ$  apart on the  $x$ -axis, while the remaining two  $p$  orbitals are perpendicular on the  $y$ -axis and the  $z$ -axis, as shown in Figure 1.22.



**Figure 1.22** An  $sp$ -hybridized carbon atom. The two  $sp$  hybrid orbitals are oriented  $180^\circ$  away from each other.

If we allow two  $sp$ -hybridized carbon atoms to approach each other,  $sp$  orbitals from each carbon can overlap head-on to form a strong  $sp$ - $sp$  sigma bond. In addition, the  $p_z$  orbitals from each carbon can form a  $p_z$ - $p_z$  pi bond by sideways overlap, and the  $p_y$  orbitals can overlap similarly to form a  $p_y$ - $p_y$  pi bond. The net effect is formation of one sigma bond and two pi bonds—a net carbon-carbon triple bond. The remaining  $sp$  hybrid orbitals each form a sigma bond with hydrogen to complete the acetylene molecule (Figure 1.23).

Because of  $sp$  hybridization, acetylene is a linear molecule with  $H-C-C$  bond angles of  $180^\circ$ . The carbon-hydrogen bond in acetylene has a length of  $1.06 \text{ \AA}$  and a strength of  $125 \text{ kcal/mol}$  ( $523 \text{ kJ/mol}$ ). The carbon-carbon



**Figure 1.23** The structure of acetylene.

bond length is  $1.20 \text{ \AA}$ , and its strength is  $200 \text{ kcal/mol}$  ( $837 \text{ kJ/mol}$ ), making the triple bond in acetylene the shortest and strongest of any carbon-carbon bond. A comparison of  $sp$ ,  $sp^2$ , and  $sp^3$  hybridization is given in Table 1.6.

**TABLE 1.6** Comparison of Carbon-Carbon and Carbon-Hydrogen Bonds in Methane, Ethane, Ethylene, and Acetylene

Molecule	Bond	Bond strength (kcal/mol)	Bond length ( $\text{\AA}$ )
Methane, $\text{CH}_4$	$\text{C}_{sp^3}-\text{H}_{1s}$	104	1.10
Ethane, $\text{CH}_3\text{CH}_3$	$\text{C}_{sp^3}-\text{C}_{sp^3}$	88	1.54
	$\text{C}_{sp^3}-\text{H}_{1s}$	98	1.10
Ethylene, $\text{H}_2\text{C}=\text{CH}_2$	$\text{C}_{sp^2}=\text{C}_{sp^2}$	152	1.33
	$\text{C}_{sp^2}-\text{H}_{1s}$	107	1.076
Acetylene, $\text{HC}\equiv\text{CH}$	$\text{C}_{sp}\equiv\text{C}_{sp}$	200	1.20
	$\text{C}_{sp}-\text{H}_{1s}$	131	1.06

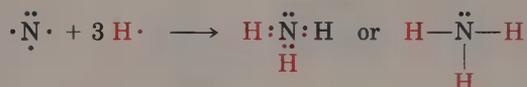
**PROBLEM** .....

- 1.14 Draw a line-bond structure for propyne,  $\text{CH}_3\text{C}\equiv\text{CH}$ . Indicate the hybridization of each carbon, and predict a value for each bond angle.
- .....

## 1.12 Hybridization of Other Atoms: Nitrogen

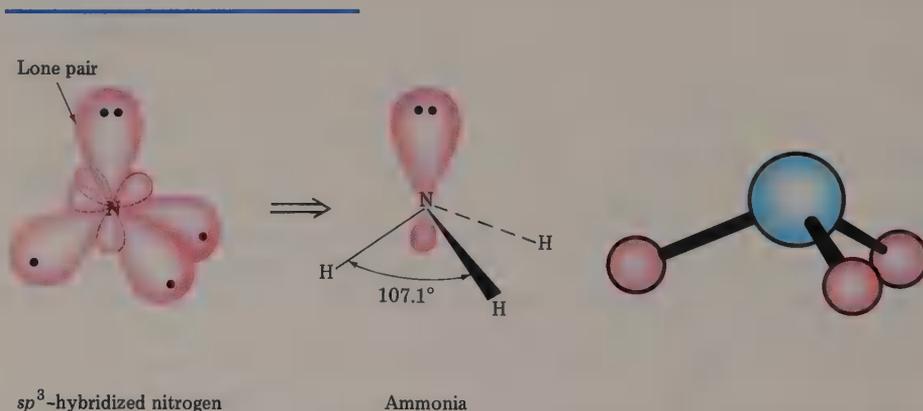
The description of covalent bonding we've developed up to this point isn't restricted to just carbon compounds. Covalent bonds formed by other elements in the periodic table also can be described in terms of hybrid orbitals. The situation becomes more complex when elements heavier than carbon are involved, but the general principles remain the same.

Let's look at ammonia,  $\text{NH}_3$ , as an example of covalent bonding involving nitrogen. A nitrogen atom has the ground-state electronic configuration  $1s^2 2s^2 2p_x 2p_y 2p_z$ , and we might therefore expect nitrogen to combine with three hydrogen atoms to complete its valence electron octet:



What geometry might ammonia have? Since the three unpaired electrons of nitrogen occupy half-filled  $2p$  orbitals, one possibility is that hydrogen  $1s$  orbitals might overlap the three nitrogen  $2p$  orbitals to form three sigma bonds. If this occurred, ammonia would have H–N–H bond angles of  $90^\circ$ . [Remember: The  $2p$  orbitals are at right angles to each other.] In fact, this picture is wrong. The experimentally measured H–N–H bond angle in ammonia is  $107.1^\circ$ , close to the tetrahedral value of  $109.5^\circ$ .

Nitrogen hybridizes to form four  $sp^3$  orbitals, exactly as carbon does. Since nitrogen has five outer-shell electrons, one of the four  $sp^3$  orbitals is occupied by two electrons and the other three each have one electron. Sigma overlap of these three half-filled nitrogen  $sp^3$  hybrid orbitals with hydrogen  $1s$  orbitals completes the ammonia molecule (Figure 1.24). Ammonia therefore has a pyramidal geometry very similar to that of methane. The N–H bond length is  $1.01 \text{ \AA}$ , and the bond strength is  $103 \text{ kcal/mol}$  ( $431 \text{ kJ/mol}$ ).



**Figure 1.24** Hybridization of nitrogen in ammonia. The nitrogen atom is  $sp^3$  hybridized, just like the carbon atom in methane.

Ammonia is pyramidal because such a structure has the lowest energy of any alternative arrangement. The energy required to hybridize the nitrogen from the ground-state configuration to  $sp^3$  configuration is more than offset by the added strength gained by bonding to  $sp^3$  orbitals (strongly directed, good overlap) versus bonding to  $p$  orbitals (poorly directed, poor overlap).

Note that an unshared electron pair is present in nitrogen, occupying an  $sp^3$  orbital. This **lone pair** of electrons occupies just as much space as an N–H bond does and is very important in the chemistry that ammonia exhibits.

PROBLEM.....

- 1.15 What kind of N–N bond do you think is present in the nitrogen molecule,  $N_2$ ? (Both nitrogens are  $sp$  hybridized.)

PROBLEM.....

- 1.16 Draw Lewis and line-bond structures for formaldimine,  $CH_2NH$ . What kind of C–N bond is present? What is the likely hybridization of the nitrogen atom?
- .....

## 1.13 Hybridization of Other Atoms: Oxygen and Boron

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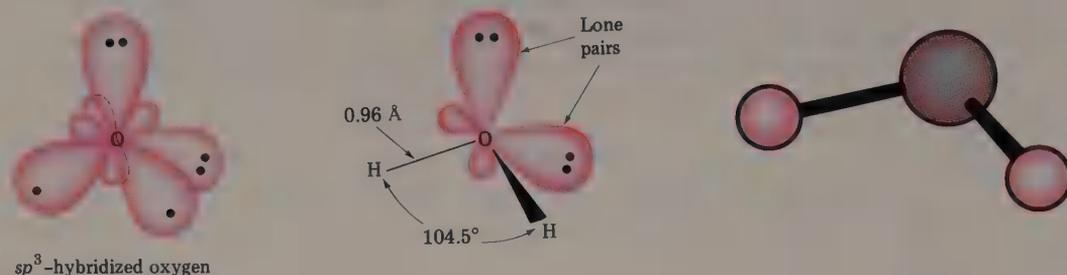
We saw in ammonia that nonbonding lone-pair electrons can occupy hybrid orbitals just as bonding electron pairs can. The same phenomenon is seen again in water,  $H_2O$ , and other oxygen-containing molecules. Ground-state oxygen has the electronic configuration  $1s^2 2s^2 2p_x^2 2p_y 2p_z$ , and oxygen is therefore divalent; that is, it forms two bonds.



We can imagine several possibilities for the bonding in water:

- Perhaps oxygen uses two unhybridized  $p$  orbitals to overlap with hydrogen  $1s$  orbitals. The two oxygen lone pairs would then occupy a  $2s$  and a  $2p$  orbital.
- Perhaps oxygen undergoes  $sp$  hybridization and uses the two  $sp$  hybrid orbitals for bonding. The lone pairs would then both occupy the two unhybridized  $p$  orbitals.
- Perhaps oxygen undergoes  $sp^2$  hybridization and uses two of the hybrid orbitals for bonding. The lone pairs would then occupy an  $sp^2$  and a  $2p$  orbital.
- Perhaps oxygen undergoes  $sp^3$  hybridization and uses two  $sp^3$  hybrid orbitals for bonding. The lone pairs would then occupy the remaining two  $sp^3$  orbitals.

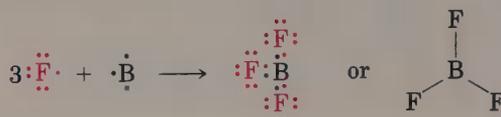
Only the fourth model, the hybridization of oxygen into  $sp^3$  orbitals, allows strong bonds and maximum distance between the valence electrons. The oxygen in water is therefore  $sp^3$  hybridized, as illustrated in Figure 1.25.



**Figure 1.25** The structure of water. The oxygen atom is  $sp^3$  hybridized like the carbon atom in methane.

Experiments on water indicate that the oxygen doesn't have perfect  $sp^3$  hybrid orbitals; the actual H-O-H bond angle of  $104.5^\circ$  is somewhat less than the predicted tetrahedral angle. We can explain this bond angle difference by assuming that there is a repulsive interaction between the two lone pairs that forces them apart, thereby compressing the H-O-H angle.

A final example of orbital hybridization that we'll consider is found in molecules such as boron trifluoride,  $BF_3$ . Since boron has only three valence electrons ( $1s^2 2s^2 2p_x$ ), it can form a maximum of three bonds. Even though we can promote a  $2s$  electron into a  $2p_y$  orbital and then hybridize in some manner, there is no way to complete a stable valence electron octet for boron.

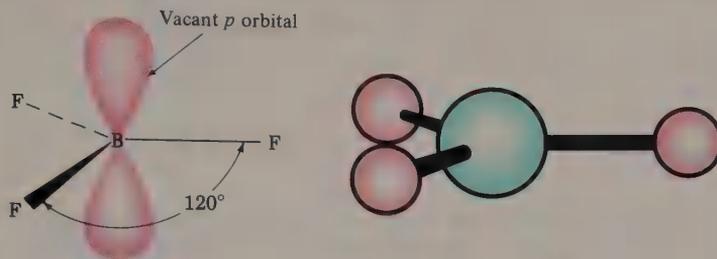


Since boron has no lone-pair electrons, we might predict that it will hybridize in such a way that the three B-F bonds will be as far away from one another as possible. This prediction implies  $sp^2$  hybridization and a planar structure for  $BF_3$  in which each fluorine bonds to a boron  $sp^2$  orbital, with the remaining  $p$  orbital on boron left vacant. Boron trifluoride has exactly this structure (Figure 1.26).

**PROBLEM** .....

**1.17** What geometry would you expect for each of the following?

- (a) The oxygen atom in methanol,  $\text{H}_3\text{C}-\ddot{\text{O}}-\text{H}$
- (b) The nitrogen atom in trimethylamine,  $\text{H}_3\text{C}-\ddot{\text{N}}-\text{CH}_3$   
 $\quad \quad \quad |$   
 $\quad \quad \quad \text{CH}_3$
- (c) The phosphorus atom in  $\text{PH}_3$
- .....



**Figure 1.26** The structure of boron trifluoride. The boron atom is  $sp^2$  hybridized and has a vacant  $p$  orbital perpendicular to the  $BF_3$  plane.

## 1.14 Summary and Key Words

**Organic chemistry** is the study of carbon compounds. Although a division into organic and inorganic chemistry occurred historically, there is no scientific reason for the division.

An atom is composed of a positively charged nucleus surrounded by one or more negatively charged electrons. The electronic structure of an atom can be described by the Schrödinger **wave equation**, in which electrons are considered to occupy **orbitals** centered around the nucleus. Different orbitals have different energy levels and different shapes. For example,  $s$  orbitals are spherical and  $p$  orbitals are dumbbell shaped. The **electronic configuration** of an atom can be found by assigning electrons to the proper orbitals, beginning with the lowest-energy ones.

There are two kinds of chemical bonds: **ionic bonds** and **covalent bonds**. Ionic bonds are based on the electrostatic attraction of unlike charges and are commonly found in inorganic salts. Covalent bonds are formed when an electron pair is shared between two atoms. Electron sharing occurs by **overlap** of two atomic orbitals to give a new **molecular orbital**. Bonds that have a circular cross-section and are formed by head-on overlap are called **sigma ( $\sigma$ ) bonds**; bonds formed by sideways overlap of two  $p$  orbitals are called **pi ( $\pi$ ) bonds**.

To form bonds in organic molecules, carbon uses hybrid orbitals. When forming only single bonds, carbon is  $sp^3$  hybridized and has four equivalent  **$sp^3$  hybrid orbitals** with tetrahedral geometry. When forming double bonds, carbon is  $sp^2$  hybridized, has three equivalent  **$sp^2$  hybrid orbitals** with planar geometry, and has one unhybridized  $p$  orbital. A carbon-carbon double bond is formed when two  $sp^2$ -hybridized carbon atoms bond together. When forming triple bonds, carbon is  $sp$  hybridized, has two equivalent  **$sp$  hybrid orbitals** with linear geometry, and has two unhybridized  $p$  orbitals. A carbon-carbon triple bond results when two  $sp$ -hybridized carbon atoms bond together.

Other atoms such as nitrogen, oxygen, and boron also hybridize in order to form stronger bonds. The nitrogen atom in ammonia and the oxygen atom in water are  $sp^3$  hybridized; the boron atom in boron trifluoride is  $sp^2$  hybridized.

## WORKING PROBLEMS

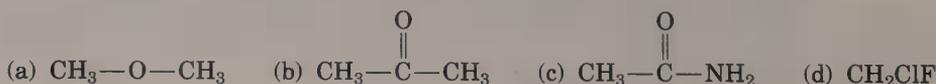
There's no surer way to learn organic chemistry than by working problems. Although careful reading and rereading of this text is important, reading alone isn't enough. In addition, you must be able to work with the information you've read and be able to use your knowledge in new situations. Working problems gives you practice at doing this.

Each chapter in this book provides many problems of different sorts. The in-chapter problems are placed for immediate reinforcement of ideas just learned. The end-of-chapter problems provide additional practice and are of two types: drill and thought. Early problems are primarily of the drill type, providing an opportunity for you to practice your command of the fundamentals. Later problems tend to be more thought provoking, and many are challenges to your depth of understanding.

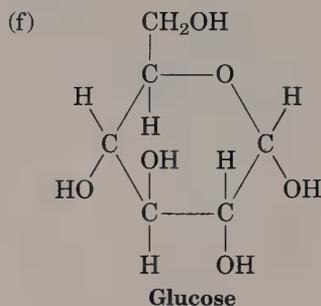
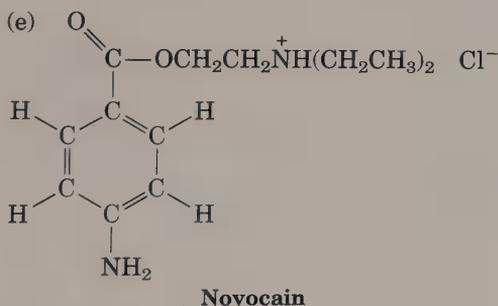
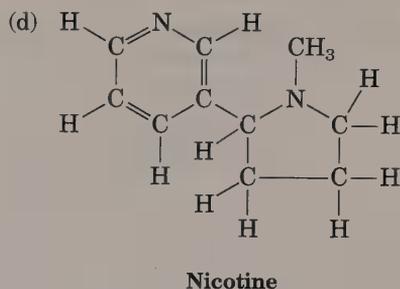
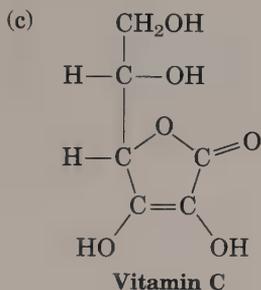
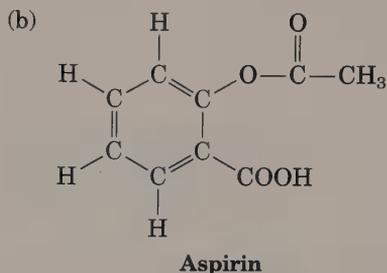
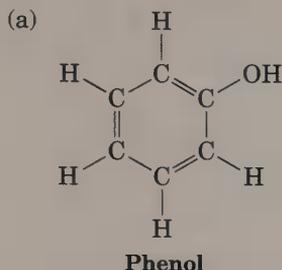
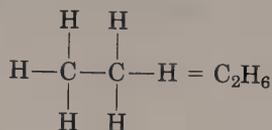
As you study organic chemistry, take the time to work the problems. Do the ones you can, and ask for help on the ones you can't do. If you're stumped by a particular exercise, check the accompanying *Study Guide and Solutions Manual* for an explanation that will help clarify the difficulty. Working problems takes effort, but the payoff in knowledge and understanding is immense.

## ADDITIONAL PROBLEMS .....

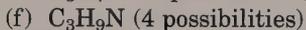
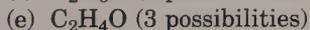
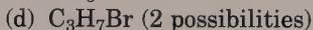
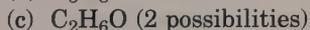
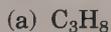
- 1.18 How many valence electrons does each of these atoms have?  
 (a) Magnesium (b) Sulfur (c) Bromine
- 1.19 Give the ground-state electronic configurations for the following elements. For example, carbon is  $1s^2 2s^2 2p^2$ .  
 (a) Sodium (b) Aluminum (c) Silicon (d) Calcium
- 1.20 What are likely formulas for these molecules?  
 (a)  $AlCl_7$  (b)  $CF_2Cl_7$  (c)  $NI_7$
- 1.21 Write Lewis (electron-dot) structures for these molecules:  
 (a)  $H-C\equiv C-H$  (b)  $AlH_3$  (c)  $CH_3-\overset{\cdot\cdot}{S}-CH_3$   
 (d)  $H_2C=CH\overset{\cdot\cdot}{C}l$  (e)  $H_2C=CH-CH=CH_2$  (f)  $CH_3-\overset{\cdot\cdot}{C}=\overset{\cdot\cdot}{O}-\overset{\cdot\cdot}{O}-H$
- 1.22 Write a Lewis structure for acetonitrile,  $H_3C-C\equiv N$ . How many electrons does the nitrogen atom have in its outer shell? How many are bonding, and how many are nonbonding?
- 1.23 Fill in any unshared electrons that are missing from the following line-bond structures:



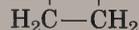
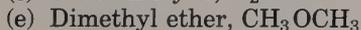
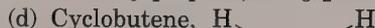
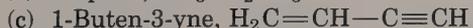
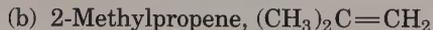
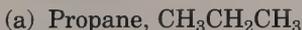
1.24 Convert the following line-bond structures into molecular formulas. For example:



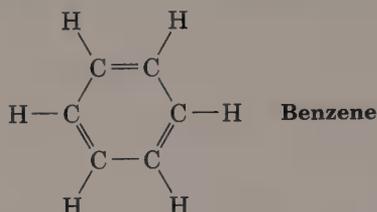
1.25 Convert the following molecular formulas into line-bond structures that are consistent with valence rules:



1.26 Indicate the kind of hybridization you would expect for each carbon atom in these molecules:

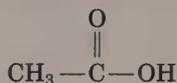


- 1.27 What is the hybridization of each carbon atom in benzene? What overall shape would you expect benzene to have?

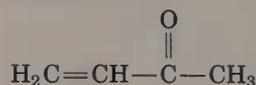


- 1.28 What kind of hybridization would you expect for the following?
- The oxygen in dimethyl ether,  $\text{CH}_3\text{-O-CH}_3$
  - The nitrogen in dimethylamine,  $\text{CH}_3\text{NHCH}_3$
  - The boron in trimethylborane,  $(\text{CH}_3)_3\text{B}$
- 1.29 On the basis of your answers to Problem 1.28, what bond angles would you expect for the following?
- The  $\text{C-O-C}$  angle in  $\text{CH}_3\text{-O-CH}_3$
  - The  $\text{C-N-C}$  angle in  $\text{CH}_3\text{NHCH}_3$
  - The  $\text{C-N-H}$  angle in  $\text{CH}_3\text{NHCH}_3$
  - The  $\text{C-B-C}$  angle in  $(\text{CH}_3)_3\text{B}$
- 1.30 What shape would you expect these species to have?
- The ammonium ion,  $\text{NH}_4^+$
  - Trimethylborane,  $(\text{CH}_3)_3\text{B}$
  - Trimethylphosphine  $(\text{CH}_3)_3\text{P}$
  - Formaldehyde,  $\text{H}_2\text{C=O}$
- 1.31 Draw a three-dimensional representation of the oxygen-bearing carbon atom in ethanol,  $\text{CH}_3\text{CH}_2\text{OH}$ , using the standard convention of solid, wedged, and dashed lines.
- 1.32 Consider the molecules  $\text{SO}_2$  and  $\text{SO}_3$  and the ion  $\text{SO}_4^{2-}$ .
- Write Lewis structures for each.
  - Predict the shape of each.
- 1.33 Draw line-bond structures for these molecules.
- Acrylonitrile,  $\text{H}_2\text{C=CHCN}$
  - Ethanol,  $\text{CH}_3\text{CH}_2\text{OH}$
  - Butane,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$
- 1.34 Sodium methoxide,  $\text{NaOCH}_3$ , contains both covalent and ionic bonds. Which do you think is which?
- 1.35 Indicate the kind of hybridization you would expect for each carbon atom in these molecules:

- (a) Acetic acid,



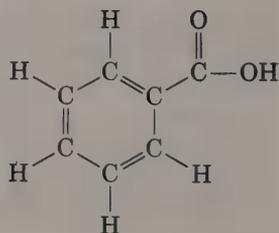
- (b) 3-Buten-2-one,



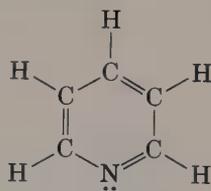
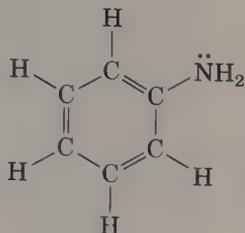
- (c) Acrylonitrile,



- (d) Benzoic acid,

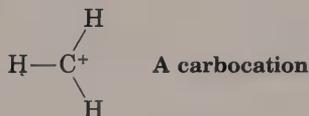


- 1.36 What kind of hybridization would you expect for the following?  
 (a) The nitrogen in aniline, (b) The nitrogen in pyridine,



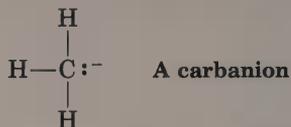
- (c) The beryllium in dimethylberyllium,  
 $\text{CH}_3\text{—Be—CH}_3$
- (d) The phosphorus in trimethylphosphine,  
 $(\text{CH}_3)_3\text{P}$

- 1.37 On the basis of your answers to Problem 1.36, what bond angles do you expect for the following?  
 (a) The C–N–H angle in aniline  
 (b) The C–Be–C angle in  $(\text{CH}_3)_2\text{Be}$   
 (c) The C–P–C angle in  $(\text{CH}_3)_3\text{P}$
- 1.38 Identify the bonds in these compounds as either ionic or covalent:  
 (a) NaCl (b)  $\text{CH}_3\text{Cl}$   
 (c)  $\text{Cl}_2$  (d) HOCl
- 1.39 Allene is an unusual molecule that has the structure  $\text{H}_2\text{C}=\text{C}=\text{CH}_2$ . Draw an orbital picture of allene. Is the central carbon atom  $sp^2$  or  $sp$  hybridized? What about the hybridization of the terminal carbons? What shape would you predict for allene?
- 1.40 Allene (Problem 1.39) is related structurally to carbon dioxide,  $\text{CO}_2$ . Draw an orbital picture of  $\text{CO}_2$ , and identify the hybridization of carbon.
- 1.41 Although almost all stable organic species have tetravalent carbon atoms, species with trivalent carbon atoms are known to exist. *Carbocations* are one such class of compounds.  
 (a) If a neutral carbon has 8 valence electrons associated with it (2 from each of 4 bonds), how many valence electrons does the positively charged carbon atom have?  
 (b) What hybridization might you expect this carbon atom to have?  
 (c) What geometry does the carbocation have?  
 (d) What relationship do you see between a carbocation and a trivalent boron compound such as  $\text{BF}_3$ ?

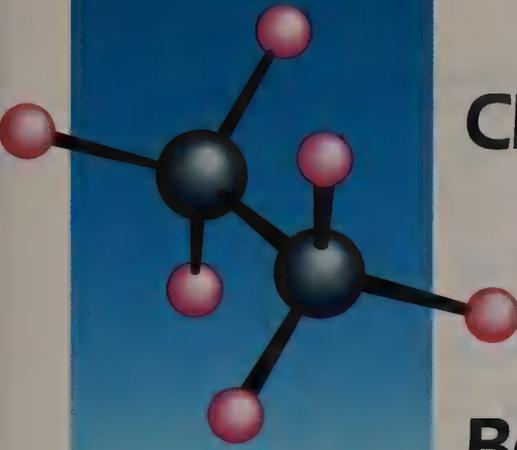


- 1.42 A *carbanion* is a species that contains a negatively charged trivalent carbon atom.  
 (a) If a neutral carbon atom has 8 valence electrons associated with it (2 from each of 4 bonds), how many valence electrons does the negatively charged carbon atom have?  
 (b) What hybridization might you expect this carbon atom to have?  
 (c) What geometry does the carbanion have?

- (d) What relationship do you see between a carbanion and a trivalent nitrogen compound such as  $\text{NH}_3$ ?



- 1.43 Divalent species called *carbenes* are known to be capable of fleeting existence. For example, methylene,  ${}^1\text{CH}_2$ , is the simplest carbene. The two unshared electrons in methylene can be either spin-paired in a single orbital or unpaired in different orbitals. Predict the type of hybridization you would expect carbon to adopt in singlet (spin-paired) methylene and triplet (spin-unpaired) methylene. Draw pictures of each, and indicate the types of carbon orbitals present.



# Chapter 2

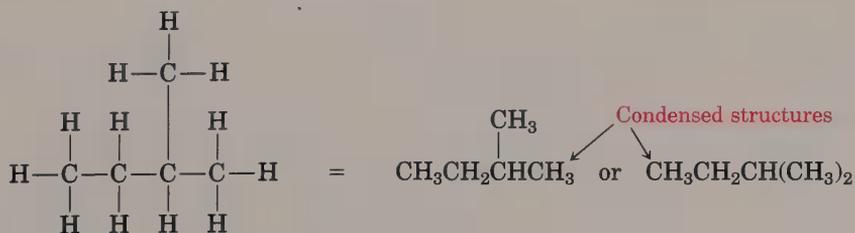
## Bonding and Molecular Properties

### 2.1 Writing and Drawing Chemical Structures

In the Kekulé structures we've been using, a line between atoms represents the two electrons in a covalent bond. Such structures have been used for many years and comprise a universal chemical language. Two chemists from different countries may not understand each other's spoken words, but a chemical structure means the same to both.

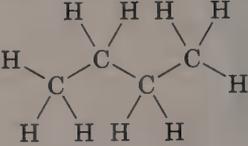
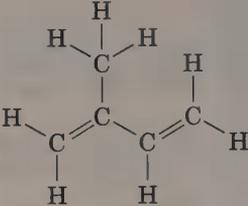
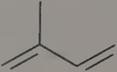
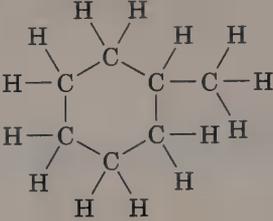
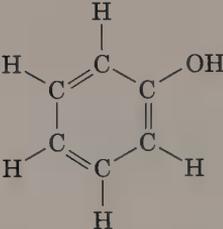
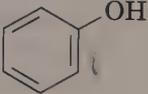
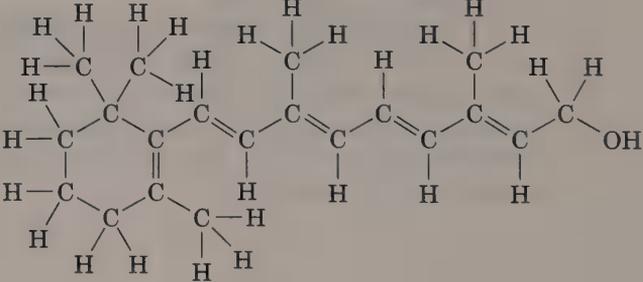
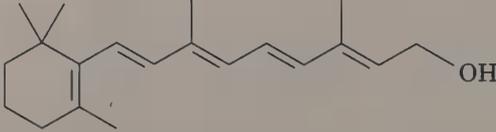
Most organic chemists find themselves drawing many structures each day, and it would soon become awkward if every bond and atom had to be indicated. For example, vitamin A,  $C_{20}H_{30}O$ , has 51 different chemical bonds uniting the 51 atoms. Vitamin A can be drawn showing each bond and atom, but doing so is a time-consuming process, and the resultant drawing is cluttered (see Table 2.1).

Chemists have therefore devised several shorthand ways of writing structures. In **condensed structures**, carbon–hydrogen and carbon–carbon bonds aren't shown, they're "understood." If a carbon has three hydrogens bonded to it, we write  $CH_3$ ; if a carbon has two hydrogens bonded to it, we write  $CH_2$ ; and so on. The compound called 2-methylbutane, for example, is written as follows:



2-Methylbutane

TABLE 2.1 Kekulé and Skeletal Structures for Several Compounds

Compound	Kekulé structure	Skeletal structure
Butane, $C_4H_{10}$		
Isoprene, $C_5H_8$		
Methylcyclohexane, $C_7H_{14}$		
Phenol, $C_6H_6O$		
Vitamin A, $C_{20}H_{30}O$		

Notice that the horizontal bonds between carbons aren't shown—the  $\text{CH}_3$ ,  $\text{CH}_2$ , and  $\text{CH}$  units are simply placed next to each other—but that the vertical carbon-carbon bond in the first condensed structure is shown for clarity. Notice also that in the second condensed structure, the two equivalent  $\text{CH}_3$  groups attached to the  $\text{CH}$  carbon are simply grouped together as  $(\text{CH}_3)_2$ .

Simpler still is the use of **skeletal structures** such as those in Table 2.1. The rules for drawing skeletal structures are simple:

1. Carbon atoms aren't usually shown. Instead, a carbon atom is assumed to be at each intersection of two lines (bonds) and at the end of each line. Occasionally, a carbon atom might be indicated for emphasis or clarity.
2. Hydrogen atoms bonded to carbon aren't shown. Since carbon always has a valence of 4, we mentally supply the correct number of hydrogen atoms to fill the valence of each carbon.
3. All atoms other than carbon and hydrogen are indicated.

Table 2.1 gives examples of how these rules are applied in specific cases.

PROBLEM .....

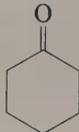
2.1 Convert these skeletal structures into molecular formulas:

(a)



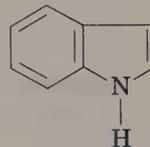
Pyridine

(b)



Cyclohexanone

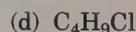
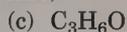
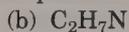
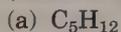
(c)



Indole

PROBLEM .....

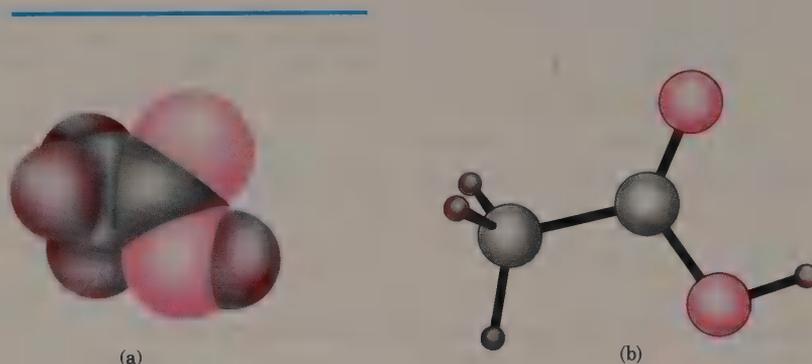
2.2 Propose skeletal structures for compounds that satisfy these molecular formulas (there is more than one possibility in each case):



## 2.2 Molecular Models

Organic chemistry is a three-dimensional science, and molecular shape often plays a crucial role in determining the chemistry a compound undergoes. One particularly helpful technique when learning organic chemistry is to use molecular models. With practice, you can learn to see many spatial relationships even when viewing two-dimensional drawings, but there's no substitute for building a molecular model and turning it in your hands to get different perspectives.

Many kinds of models are available, some at relatively modest cost, and every student should have ready access to a set of models while studying this book. Research chemists generally prefer to use either space-filling models such as Corey–Pauling–Koltun (CPK™) Molecular Models or skeletal models such as Dreiding Stereomodels™. Both are expensive but are precisely made to reflect accurate bond angles, intramolecular distances, and atomic radii. CPK models are generally preferred for examining the degree of crowding within a molecule, while skeletal models let the user measure bond angles and interatomic distances more readily. For student use, ball-and-stick models are generally the least expensive and most durable. Figure 2.1 shows two kinds of models of acetic acid,  $\text{CH}_3\text{COOH}$ .



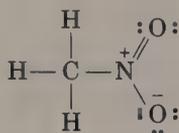
**Figure 2.1** Molecular models of acetic acid,  $\text{CH}_3\text{COOH}$ : (a) space-filling; (b) ball-and-stick.

PROBLEM .....

- 2.3 Build a molecular model of ethane,  $\text{H}_3\text{C}-\text{CH}_3$ , and look at the relationships between hydrogens on the different carbons by sighting along the C–C bond.
- .....

## 2.3 Formal Charges

Most organic molecules can be accurately represented by the Kekulé line-bond structures we saw in Chapter 1. Sometimes, though, electron book-keeping requires that we attach **formal charges** to specific atoms within a molecule. This is particularly true for atoms that have an apparently “abnormal” number of bonds. In nitromethane ( $\text{CH}_3\text{NO}_2$ ), for example, the nitrogen atom has four bonds rather than the usual three and must be represented as having a positive charge, while the single bound oxygen atom has one bond rather than the usual two and must be represented as having a negative charge:

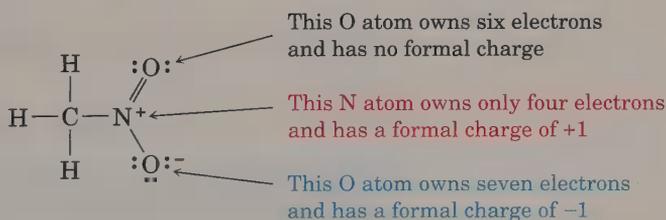


Nitromethane

Why and how are formal charges assigned? A normal covalent bond is formed when each atom donates one electron. Although the bonding electrons are shared by both atoms, each atom can still be thought of as “owning” one electron for bookkeeping purposes. In methane, for example, carbon can be considered to own one electron from each of the four bonds for a total of four. Since a neutral, isolated carbon atom has four valence electrons, and since the carbon atom in methane still owns four, the methane carbon is electrically neutral and has no formal charge.

The same is true for ammonia, which has three covalent N–H bonds. Atomic nitrogen has five valence electrons, and the ammonia nitrogen also has five (one from each of three shared N–H bonds plus two in the lone pair). Thus, the nitrogen atom in ammonia is electrically neutral.

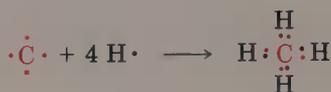
The situation is different for nitromethane. Atomic nitrogen has five valence electrons, but the nitromethane nitrogen owns only *four* (one from the C–N bond, one from the NO single bond, and two from the N=O double bond). Thus, the nitrogen has formally lost an electron and must therefore have a positive charge. A similar calculation for the singly bonded oxygen atom shows that it has formally gained an electron and must have a negative charge. (Atomic oxygen has six valence electrons, but the singly bound oxygen in nitromethane has seven—one from the O–N bond and two from each of three lone pairs.)



It's a good idea to work out the formal charges in a molecule in a logical manner to make sure you understand the reasons behind the answers. To express the calculations in a general way, we can say that every atom in a molecule can be assigned a formal charge equal to the number of valence electrons in a neutral, isolated atom minus the number of electrons still owned by that atom in the molecule:

$$\begin{aligned} \text{Formal charge} &= \left( \begin{array}{c} \text{Number of} \\ \text{valence electrons} \\ \text{in free atom} \end{array} \right) - \left( \begin{array}{c} \text{Number of} \\ \text{valence electrons} \\ \text{in bound atom} \end{array} \right) \\ &= \left( \begin{array}{c} \text{Number of} \\ \text{valence} \\ \text{electrons} \end{array} \right) - \left( \begin{array}{c} \text{Half of} \\ \text{bonding} \\ \text{electrons} \end{array} \right) - \left( \begin{array}{c} \text{Number of} \\ \text{nonbonding} \\ \text{electrons} \end{array} \right) \end{aligned}$$

For the methane carbon,



$$\text{Carbon valence electrons} = 4$$

$$\text{Carbon bonding electrons} = 8$$

$$\text{Carbon nonbonding electrons} = 0$$

$$\text{Formal charge} = 4 - \frac{8}{2} - 0 = 0$$

For the ammonia nitrogen,



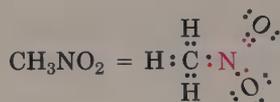
$$\text{Nitrogen valence electrons} = 5$$

$$\text{Nitrogen bonding electrons} = 6$$

$$\text{Nitrogen nonbonding electrons} = 2$$

$$\text{Formal charge} = 5 - \frac{6}{2} - 2 = 0$$

For the nitromethane nitrogen,



$$\text{Nitrogen valence electrons} = 5$$

$$\text{Nitrogen bonding electrons} = 8$$

$$\text{Nitrogen nonbonding electrons} = 0$$

$$\text{Formal charge} = 5 - \frac{8}{2} - 0 = +1$$

For the singly bound nitromethane oxygen,

$$\text{Oxygen valence electrons} = 6$$

$$\text{Oxygen bonding electrons} = 2$$

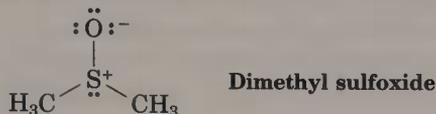
$$\text{Oxygen nonbonding electrons} = 6$$

$$\text{Formal charge} = 6 - \frac{2}{2} - 6 = -1$$

Molecules such as nitromethane that are neutral overall but have plus and minus charges on individual atoms are said to be **dipolar**. Dipolar character in molecules often has important chemical consequences, and it's important to be able to identify and calculate the charges correctly.

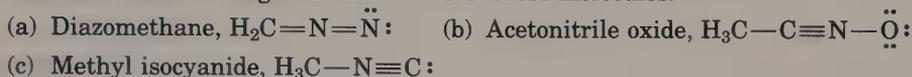
## PROBLEM .....

- 2.4 Dimethyl sulfoxide, a common solvent, has the structure indicated. Show by calculations why dimethyl sulfoxide must have formal charges on S and O.



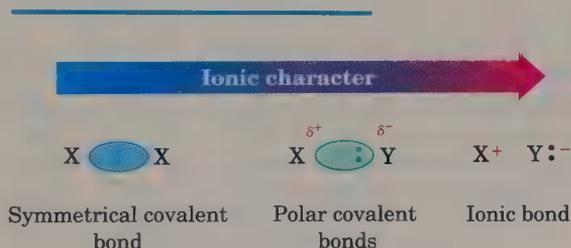
## PROBLEM .....

- 2.5 Calculate formal charges for the atoms in these molecules:



## 2.4 Polar Covalent Bonds: Electronegativity

Up to this point, we've viewed chemical bonding in an either/or manner: A given bond is either covalent or ionic. It's more accurate, though, to look at bonding as a continuum of possibilities, from a perfectly covalent bond with a symmetrical electron distribution on the one hand, to a perfectly ionic bond between positive and negative ions on the other (Figure 2.2).



**Figure 2.2** The continuum in bonding from covalent to ionic as a result of unsymmetrical electron distribution. The symbol  $\delta$  (Greek delta) means *partial charge*, either partial positive ( $\delta^+$ ) or partial negative ( $\delta^-$ ).

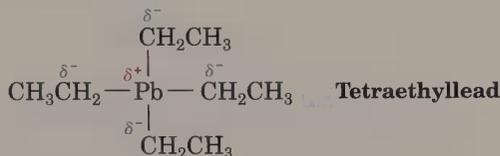
The carbon–carbon bond in ethane, for example, is electronically symmetrical and therefore perfectly covalent; the two bonding electrons are equally shared by the two equivalent carbon atoms. The bond in sodium chloride, by contrast, is largely ionic.<sup>1</sup> An electron has been donated from sodium to chlorine to give  $\text{Na}^+$  and  $\text{Cl}^-$  ions that are held together in the solid by electrostatic attraction. Between these two extremes lie the great

<sup>1</sup>Even the bond in  $\text{Na}^+ \text{Cl}^-$  is only about 80% ionic rather than 100%.



An arrow  $\leftrightarrow$  is used to indicate the direction of polarity. By convention, *electrons move in the direction of the arrow*. The tail of the arrow is electron-poor ( $\delta^+$ ), and the head of the arrow is electron-rich ( $\delta^-$ ).

Bonds between carbon and less electronegative elements are polarized so that carbon bears a partial negative charge and the other atom bears a partial positive charge. Organometallic compounds such as tetraethyllead, the “lead” in gasoline, provide good examples of this kind of polar bond:



When speaking of an atom’s ability to polarize a bond, we use the term **inductive effect**. An inductive effect is simply the shifting of electrons in a bond in response to the electronegativity of nearby atoms. Electropositive elements such as lithium and magnesium inductively donate electrons, whereas electronegative elements such as oxygen and chlorine inductively withdraw electrons. Inductive effects play a major role in understanding chemical reactivity, and we’ll use them many times throughout this text to explain a wide variety of chemical phenomena.

PROBLEM.....

- 2.6 Which element in each of the following pairs is more electronegative?  
 (a) Li or H                      (b) Be or Br                      (c) Cl or I                      (d) C or H

PROBLEM.....

- 2.7 Use the  $\delta^+/\delta^-$  convention to indicate the direction of expected polarity for each of the bonds indicated.  
 (a)  $\text{H}_3\text{C}-\text{Br}$                       (b)  $\text{H}_3\text{C}-\text{NH}_2$                       (c)  $\text{H}_3\text{C}-\text{Li}$                       (d)  $\text{H}_2\text{N}-\text{H}$   
 (e)  $\text{H}_3\text{C}-\text{OH}$                       (f)  $\text{H}_3\text{C}-\text{MgBr}$                       (g)  $\text{H}_3\text{C}-\text{F}$
- .....

## 2.5 Polar Covalent Bonds: Dipole Moment

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Since individual bonds are often polar, molecules as a whole are often polar also. Overall molecular polarity results from the summation of all individual bond polarities, formal charges, and lone-pair contributions in the molecule. The measure of this net molecular polarity is a quantity called the *dipole moment*. In practical terms, polar substances are often soluble in polar solvents like water, whereas nonpolar substances are insoluble in water.

Dipole moments can be viewed in the following way: Assume that there is a “center of mass” of all positive charges (nuclei) in a molecule and a

center of mass of all negative charges (electrons) in the molecule. If these two centers don't coincide, then the molecule is electrically unsymmetrical and has a net polarity. The **dipole moment**,  $\mu$  (Greek mu), is defined as the magnitude of a unit charge  $q$  times the distance  $r$  between the centers. When  $q$  is measured in electrostatic units (esu) and  $r$  is measured in angstroms, dipole moments are expressed in debye<sup>2</sup> units (D):

$$\mu = (q) \times (r) \times (10^{10} \text{ D/esu } \text{\AA})$$

where

$q$  = Electric charge in electrostatic units (esu)

$r$  = Distance in angstroms

For example, if one proton and one electron (charge  $q = 4.8 \times 10^{-10}$  esu) are separated from each other by 1.00  $\text{\AA}$ , then the dipole moment is 4.8 D.

$$\mu = (4.8 \times 10^{-10} \text{ esu}) \times (1.00 \text{ \AA}) \times (10^{10} \text{ D/esu} \cdot \text{\AA}) = 4.8 \text{ D}$$

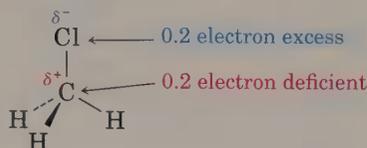
It's relatively easy to measure dipole moments experimentally, and values for some common substances are given in Table 2.2. Once the dipole moment is known, it's then possible to work backward to get an idea of the amount of charge separation in a molecule. For example, let's take chloromethane,  $\mu = 1.87$  D. If we assume that the contributions of the nonpolar C-H bonds are small, then most of the chloromethane dipole moment is due to the C-Cl bond. Since the C-Cl bond distance is 1.78  $\text{\AA}$ , the dipole moment of chloromethane would be  $1.78 \times 4.8 \text{ D} = 8.5 \text{ D}$  if the C-Cl bond were ionic. But because the measured dipole moment is only 1.87 D, the C-Cl

TABLE 2.2 Dipole Moments of Some Compounds

Compound	Dipole moment (D)	Compound	Dipole moment (D)
NaCl	9.0	NH <sub>3</sub>	1.47
	3.46	CH <sub>4</sub>	0
Nitromethane		CCl <sub>4</sub>	0
CH <sub>3</sub> Cl	1.87	CH <sub>3</sub> CH <sub>3</sub>	0
H <sub>2</sub> O	1.85		0
CH <sub>3</sub> OH	1.70	Benzene	
	1.50	BF <sub>3</sub>	0
Diazomethane			

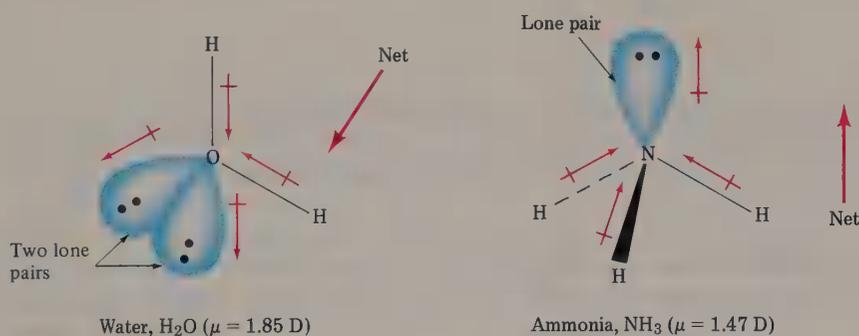
<sup>2</sup>Peter Joseph Wilhelm Debye (1884–1966); b. Maastricht, Netherlands; Ph.D. Munich (1910); professor of physics, Zurich, Utrecht, Göttingen, Leipzig, Berlin; professor of chemistry, Cornell University (1936–1966); Nobel Prize (1936).

bond is only about  $1.87/8.5 = 20\%$  ionic. Thus, the chlorine atom in chloromethane has an excess of about 0.2 electron, while the carbon atom has a deficiency of about 0.2 electron.



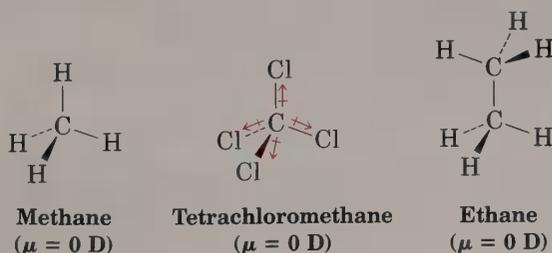
Chloromethane ( $\mu = 1.87$  D)

The two largest dipole moments listed in Table 2.2 belong to NaCl, which is highly ionic, and nitromethane ( $\text{CH}_3\text{NO}_2$ ), which has formal charges on two atoms (Section 2.3). Water and ammonia (Figure 2.4) also have relatively large dipole moments because both oxygen and nitrogen are electron-withdrawing relative to hydrogen according to their electronegativities in Figure 2.3. In addition, lone-pair electrons like those present on the oxygen atom of water and the nitrogen atom of ammonia normally make large contributions to overall dipole moments because they have no atom attached to them to “neutralize” their negative charge.



**Figure 2.4** Dipole moments of water and ammonia. The lone-pair electrons make large contributions to the dipole moments of both molecules.

By contrast with water and ammonia, methane, tetrachloromethane, and ethane have zero dipole moments. Because of their symmetrical structures, the individual bond polarities in these molecules exactly cancel each other.



## PROBLEM .....

2.8 Account for the observed dipole moments of methanol (1.70 D) and benzene (0 D) by using arrows to indicate the direction in which electron density is displaced.

(a) Methanol,  $\text{CH}_3\text{OH}$  (1.70 D)

(b) Benzene,  (0 D)

## PROBLEM .....

2.9 Carbon dioxide,  $\text{CO}_2$ , has a dipole moment of zero even though carbon–oxygen bonds are strongly polarized. Explain.

## PROBLEM .....

2.10 Make three-dimensional drawings of these compounds, and predict whether each has a dipole moment. If a dipole moment is expected, show its direction.

(a)  $\text{H}_2\text{C}=\text{CH}_2$

(b)  $\text{CHCl}_3$

(c)  $\text{CH}_2\text{Cl}_2$

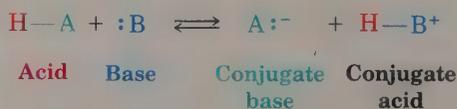
(d)  $\text{H}_2\text{C}=\text{CCl}_2$

## 2.6 Acids and Bases: The Brønsted–Lowry Definition

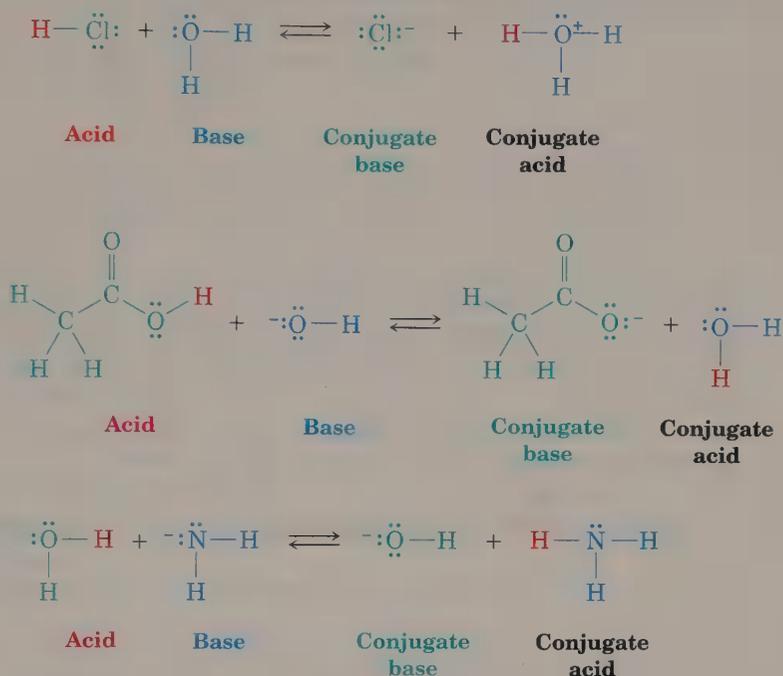
Acidity and basicity are related to the concepts of electronegativity and polarity just described, and it's a good idea at this point to review some fundamental ideas about these topics. We'll soon see that the acid–base behavior of organic molecules helps explain much of their chemistry.

According to the **Brønsted–Lowry definition**, an acid is a substance that donates a proton (hydrogen ion,  $\text{H}^+$ ), and a base is a substance that accepts a proton. When  $\text{HCl}$  gas dissolves in water, for example, an acid–base reaction occurs. Hydrogen chloride acts as an acid to donate a proton, and water acts as a base to accept the proton. The products of the reaction are  $\text{H}_3\text{O}^+$  and  $\text{Cl}^-$ . Chloride ion ( $\text{Cl}^-$ ), the product that results when the acid  $\text{HCl}$  loses a proton, is called the **conjugate base** of the acid.  $\text{H}_3\text{O}^+$ , the product that results when the base  $\text{H}_2\text{O}$  gains a proton, is called the **conjugate acid** of the base. Other common mineral acids such as  $\text{H}_2\text{SO}_4$ ,  $\text{HNO}_3$ , and  $\text{HBr}$  behave similarly, as do organic carboxylic acids such as acetic acid,  $\text{CH}_3\text{COOH}$  (Section 20.3).

In a general sense:

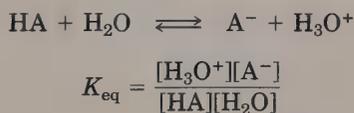


For example:



Note in the above examples that water can act *either* as an acid or as a base. In its reaction with HCl, water is a base that accepts a proton to give the hydronium ion,  $\text{H}_3\text{O}^+$ , but in its reaction with amide ion,  $^-\text{NH}_2$ , water is an acid that donates a proton to give ammonia,  $\text{NH}_3$ , and hydroxide ion,  $\text{HO}^-$ .

Acids differ in their proton-donating ability. Stronger acids such as HCl react almost completely with water, whereas weaker acids such as acetic acid ( $\text{CH}_3\text{COOH}$ ) react only slightly. The exact strength of a given acid in water solution can be described using the standard equilibrium expression for equilibrium constant,  $K_{\text{eq}}$ :

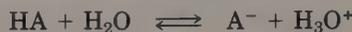


where HA represents any acid.<sup>3</sup>

In the dilute aqueous solution normally used for measuring  $K_{\text{eq}}$ , the concentration of water,  $[\text{H}_2\text{O}]$ , remains nearly constant at approximately 55.5 M. We can therefore rewrite the equilibrium expression using a new

<sup>3</sup>Remember that brackets, [ ], refer to the concentration of the enclosed species in moles per liter.

term called the **acidity constant**,  $K_a$ . The acidity constant for any generalized acid, HA, is simply the equilibrium constant multiplied by the molar concentration of pure water, 55.5 M:



$$K_a = K_{\text{eq}}[\text{H}_2\text{O}] = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]}$$

Stronger acids have their equilibria toward the right and thus have larger acidity constants, whereas weaker acids have their equilibria toward the left and have smaller acidity constants. The range of  $K_a$  values for different acids is enormous, running from about  $10^{15}$  for the strongest inorganic acids to about  $10^{-60}$  for the weakest organic acids such as methane. The common inorganic acids such as  $\text{H}_2\text{SO}_4$ ,  $\text{HNO}_3$ , and  $\text{HCl}$  have  $K_a$ 's in the range  $10^2$ – $10^9$ , while organic acids have  $K_a$ 's in the range  $10^{-5}$ – $10^{-15}$ . As you gain more experience in later chapters, you'll develop a rough feeling for which acids are "strong" and which are "weak."

Acid strengths are normally expressed by quoting  $\text{p}K_a$  values, where the  $\text{p}K_a$  is equal to the negative logarithm of the acidity constant:

$$\text{p}K_a = -\log K_a$$

A stronger acid (larger acidity constant,  $K_a$ ) has a *lower*  $\text{p}K_a$ , and a weaker acid (smaller  $K_a$ ) has a *higher*  $\text{p}K_a$ . Table 2.3 lists the  $\text{p}K_a$ 's of some common acids in order of their strength. A much more comprehensive table is given in Appendix B.

**TABLE 2.3** Relative Strength of Some Common Acids and Their Conjugate Bases

	Acid	Name	$\text{p}K_a$	Conjugate base	Name			
Weaker acid	$\text{CH}_3\text{CH}_2\text{OH}$	Ethanol	16.00	$\text{CH}_3\text{CH}_2\text{O}^-$	Ethoxide ion	Stronger base		
	$\text{H}_2\text{O}$	Water	15.74	$\text{HO}^-$	Hydroxide ion			
	$\text{HCN}$	Hydrocyanic acid	9.2	$\text{CN}^-$	Cyanide ion			
	$\text{CH}_3\text{COOH}$	Acetic acid	4.72	$\text{CH}_3\text{COO}^-$	Acetate ion			
	$\text{HF}$	Hydrofluoric acid	3.2	$\text{F}^-$	Fluoride ion			
	$\text{HNO}_3$	Nitric acid	-1.3	$\text{NO}_3^-$	Nitrate ion			
	Stronger acid	$\text{HCl}$	Hydrochloric acid	-7.0	$\text{Cl}^-$		Chloride ion	Weaker base

Notice that there is an inverse relationship shown in Table 2.3 between the acid strength of an acid and the base strength of its conjugate base. The conjugate base of a *stronger* acid is a *weaker* base because it has little affinity for a proton. Similarly, the conjugate base of a weaker acid is a stronger base because it has a high affinity for a proton. For example, chloride ion,

the conjugate base of the stronger acid HCl, is a weak base because it has little affinity for a proton; acetate ion, the conjugate base of the weaker acid  $\text{CH}_3\text{COOH}$ , is a stronger base with a modest affinity for a proton; and hydroxide ion, the conjugate base of the weaker acid  $\text{H}_2\text{O}$ , is a still stronger base with a high affinity for a proton.

PROBLEM.....

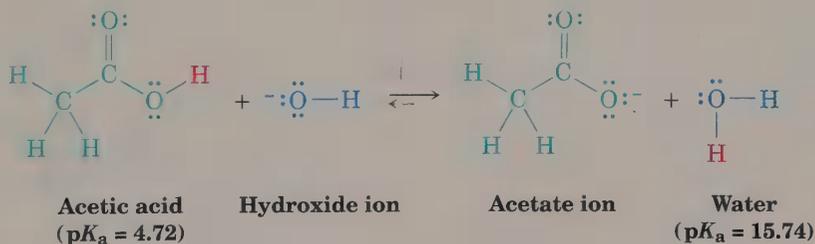
- 2.11 Formic acid,  $\text{HCOOH}$ , has  $pK_a = 3.7$ , and picric acid,  $\text{C}_6\text{H}_3\text{N}_3\text{O}_7$ , has  $pK_a = 0.3$ . Which is the stronger acid?

PROBLEM.....

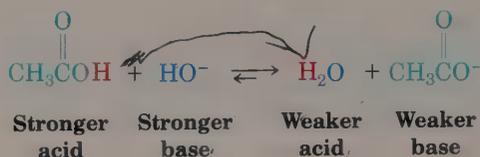
- 2.12 Amide ion,  $\text{H}_2\text{N}^-$ , is a much stronger base than hydroxide ion,  $\text{HO}^-$ . Which would you expect to be a stronger acid,  $\text{NH}_3$  or  $\text{H}_2\text{O}$ ? Explain.

## 2.7 Predicting Acid–Base Reactions from $pK_a$ Values

Compilations of  $pK_a$  values like those in Table 2.3 are extremely useful for predicting whether or not a given acid–base reaction will take place. In general, an acid will donate a proton to the conjugate base of any acid with a higher  $pK_a$ . Conversely, the conjugate base of an acid will abstract a proton from any acid with a lower  $pK_a$ . For example, the data in Table 2.3 indicate that hydroxide ion will react with acetic acid,  $\text{CH}_3\text{COOH}$ , to yield acetate ion,  $\text{CH}_3\text{COO}^-$ , and water. Since water ( $pK_a = 15.74$ ) is a weaker acid than acetic acid ( $pK_a = 4.72$ ) is, hydroxide ion has a greater affinity for a proton than acetate ion has, and the reaction of hydroxide ion with acetic acid will occur:

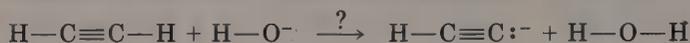


Another way for predicting acid–base reactivity is to remember that the products must be more stable than the reactants in order for reaction to occur. In other words, the product acid and base must be weaker and less reactive than the starting acid and base. For example, in the reaction of acetic acid with hydroxide ion, the product conjugate base (acetate ion) is weaker than the starting base (hydroxide ion), and the product conjugate acid (water) is weaker than the starting acid (acetic acid):



## PRACTICE PROBLEM.....

Water has  $\text{p}K_a = 15.74$ , and acetylene has  $\text{p}K_a = 25$ . Which of the two is more acidic? Would you expect hydroxide ion to react with acetylene?



**Solution** In comparing two acids, the one with the lower  $\text{p}K_a$  is stronger. Thus, water is a stronger acid than acetylene. Since water gives up a proton more easily than acetylene does, the  $\text{HO}^-$  ion must have less affinity for a proton than the  $\text{HC}\equiv\text{C}:^-$  ion has. In other words, the anion of acetylene is a stronger base than hydroxide ion, and the reaction will not proceed as written.

## PRACTICE PROBLEM.....

According to the data in Table 2.3, acetic acid has  $\text{p}K_a = 4.72$ . What is its  $K_a$ ?

**Solution** Since  $\text{p}K_a$  is the negative logarithm of  $K_a$ , it's necessary to use a calculator with an ANTILOG or INV LOG function. Change the sign and enter the value of the  $\text{p}K_a$  ( $-4.72$ ) and then find its antilog ( $1.91 \times 10^{-5}$ ). Thus,  $K_a = 1.91 \times 10^{-5}$  for HCN.

## PRACTICE PROBLEM.....

Use the  $K_a$  value for acetic acid calculated in the preceding Practice Problem to find the concentration of  $\text{H}_3\text{O}^+$  in a 0.100 M solution of acetic acid.

**Solution** Let the concentration of  $\text{H}_3\text{O}^+$  be represented by  $[X]$ . The concentration of acetate ion is also  $[X]$  since there is a 1:1 ratio of  $\text{H}_3\text{O}^+$  and acetate, and the concentration of acetic acid is  $[0.100 - X]$ . But since acetic acid is a weak acid, we know that the value of  $X$  is small, and we can assume that  $[0.100 - X]$  is approximately 0.100. The equilibrium equation then becomes

$$\begin{aligned}
 K_a &= \frac{[\text{H}_3\text{O}^+][\text{Acetate}]}{[\text{Acetic acid}]} \\
 1.91 \times 10^{-5} &= \frac{[X][X]}{[0.100 - X]} = \frac{[X]^2}{[0.100 - X]} \approx \frac{[X]^2}{0.100} \\
 \text{so } [X]^2 &= 1.91 \times 10^{-6}, \text{ and } X = 1.38 \times 10^{-3} \text{ M}
 \end{aligned}$$

The  $\text{H}_3\text{O}^+$  concentration is therefore  $1.38 \times 10^{-3}$  M, corresponding to  $\text{pH} = 2.9$ .

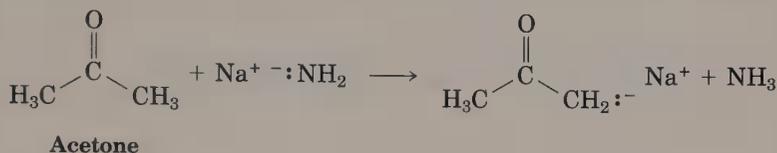
## PROBLEM.....

**2.13** Is either of the following reactions likely to take place, according to the  $\text{p}K_a$  data in Table 2.3?

- (a)  $\text{HCN} + \text{CH}_3\text{COO}^- \text{Na}^+ \longrightarrow \text{Na}^+ \text{CN}^- + \text{CH}_3\text{COOH}$   
 (b)  $\text{CH}_3\text{CH}_2\text{OH} + \text{Na}^+ \text{CN}^- \longrightarrow \text{CH}_3\text{CH}_2\text{O}^- \text{Na}^+ + \text{HCN}$

## PROBLEM.....

- 2.14 Ammonia,  $\text{NH}_3$ , has  $\text{p}K_a \approx 36$  and acetone has  $\text{p}K_a \approx 20$ . Will the following reaction take place?



## PROBLEM.....

- 2.15 What is the  $K_a$  of HCN if its  $\text{p}K_a$  is 9.2?

## PROBLEM.....

- 2.16 What is the pH of a 0.500 M aqueous solution of benzoic acid,  $\text{p}K_a = 4.2$ ?

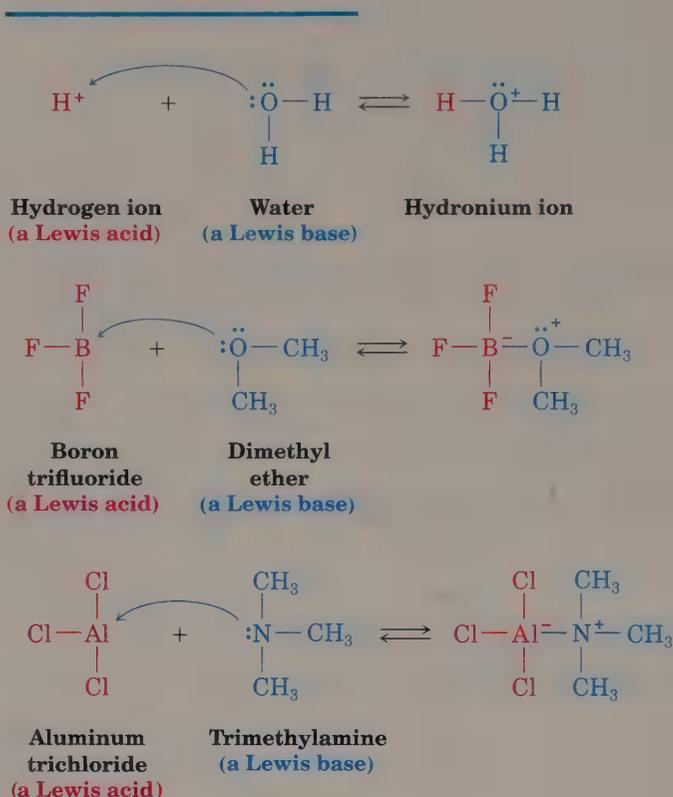
## 2.8 Acids and Bases: The Lewis Definition

The Brønsted–Lowry definition of acidity discussed in the previous two sections is an extremely useful one that can be extended to all compounds containing hydrogen. Of even more use, however, is the *Lewis definition* of acids and bases. A **Lewis acid** is any substance that accepts an electron pair; a **Lewis base** is any substance that donates an electron pair in forming a covalent bond.

### LEWIS ACIDS

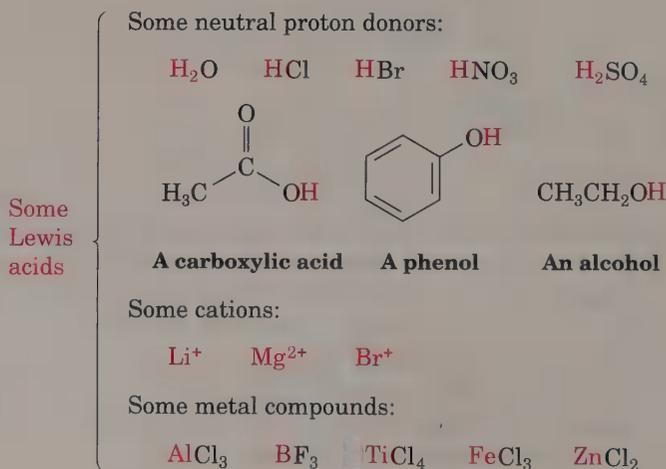
The fact that a Lewis acid must be able to accept an electron pair means that it must have a vacant, low-energy orbital it can use for bonding. As a result, the Lewis definition of acidity is much broader than the Brønsted–Lowry definition. It includes not only  $\text{H}^+$  donors but many other species as well. For example, positively charged ions including the proton (hydrogen ion,  $\text{H}^+$ ) and various metal cations such as  $\text{Mg}^{2+}$  are Lewis acids because they accept a pair of electrons when they form a bond to a base. In the same way, compounds of Group 3A elements such as  $\text{BF}_3$  and  $\text{AlCl}_3$  are Lewis acids because they have unfilled valence shells and can accept electron pairs from Lewis bases, as shown in Figure 2.5. Similarly, many transition-metal compounds such as  $\text{TiCl}_4$ ,  $\text{FeCl}_3$ ,  $\text{ZnCl}_2$ , and  $\text{SnCl}_4$  are Lewis acids.

Note how the acid–base reactions in Figure 2.5 are shown. A curved arrow indicates the direction of electron-pair flow from the electron-rich Lewis base to the electron-poor Lewis acid. This kind of arrow is used extensively in organic chemistry and always has the same meaning: A pair of electrons moves from the atom at the tail of the arrow to form a bond with the atom at the head of the arrow. *An arrow represents the movement of electrons, not the movement of atoms.*



**Figure 2.5** The reactions of some Lewis acids with some Lewis bases. The Lewis acids all have a vacant orbital that can accept an electron pair; the Lewis bases all have a pair of nonbonding electrons. Note how the flow of electrons from the Lewis base to the Lewis acid is indicated by the curved arrows.

Some further examples of Lewis acids are shown below.









the tube before and after combustion, we can accurately determine the amount of water formed and then calculate the amount of hydrogen present in the original sample. The  $\text{CO}_2$  produced by combustion passes through the  $\text{CaCl}_2$  tube into a separate tube containing potassium hydroxide,  $\text{KOH}$ , where it is absorbed. Again, the amount of  $\text{CO}_2$  present can be determined by weighing the tube before and after combustion, and the percentage of carbon present in the original sample can be determined. The Liebig technique can't directly determine the amount of oxygen present in a compound. If, however, no other elements are detected and the combined percentages of carbon and hydrogen don't total 100, then the percentage of oxygen present is taken as the difference.

Let's assume that we've analyzed a 0.550 gram (g) sample of a colorless organic liquid obtained by the distillation of wine. On weighing the  $\text{CaCl}_2$  and  $\text{KOH}$  tubes, we find that 0.660 g of  $\text{H}_2\text{O}$  and 1.037 g of  $\text{CO}_2$  have been formed. We can then calculate the percentages of carbon and hydrogen in the unknown sample.

First, we find the weight of hydrogen in the sample by finding how much water is produced. We then calculate the percent hydrogen in the sample by dividing the hydrogen weight by the sample weight:

$$\begin{aligned}\text{Weight of H in sample} &= \text{Weight of H}_2\text{O} \times \frac{\text{Molecular weight of H}_2 (2.016)}{\text{Molecular weight of H}_2\text{O} (18.016)} \\ &= (0.66)(0.112) = 0.074 \text{ g H}\end{aligned}$$

$$\% \text{ H in sample} = \frac{\text{Weight of H}}{\text{Weight of sample}} = \frac{0.074}{0.55} = 13.44\%$$

In a similar manner, we calculate the percent carbon in the sample by first finding the weight of carbon from the amount of  $\text{CO}_2$  produced and then dividing by the sample weight:

$$\begin{aligned}\text{Weight of C in sample} &= \text{Weight of CO}_2 \times \frac{\text{Molecular weight of C} (12.01)}{\text{Molecular weight of CO}_2 (44.01)} \\ &= (1.037)(0.273) = 0.283 \text{ g C}\end{aligned}$$

$$\% \text{ C in sample} = \frac{\text{Weight of C}}{\text{Weight of sample}} = \frac{0.283}{0.55} = 51.47\%$$

Since the percentages of carbon and hydrogen add up to only 64.91%, we can assume that the sample also contains 35.09% oxygen. The next step is to determine the atomic ratios by dividing the percentage of each element by its atomic weight:

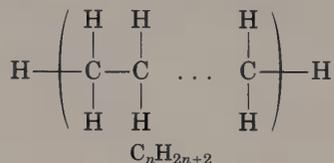
$$\text{C: } \frac{51.47\%}{12.01} = 4.29$$

$$\text{H: } \frac{13.44\%}{1.008} = 13.33$$

$$\text{O: } \frac{35.09\%}{16.00} = 2.19$$

The atomic ratio of elements in the sample is C, 4.28:H, 13.33:O, 2.193, which reduces to C, 1.95:H, 6.1:O, 1 when the numbers are divided by 2.193 (the lowest number in the series). Rounding off these ratios gives the empirical formula  $C_2H_6O$ .

It's important to remember that analysis gives only atomic ratios. To determine the molecular formula, which may be a multiple of the empirical formula, we also need to determine the molecular weight. In the present case, though, any higher multiple of  $C_2H_6O$  would be impossible by the rules of valency. Since carbon has a valence of 4, an organic compound with  $n$  carbons can have no more than  $2n + 2$  hydrogens—that is, two hydrogen atoms per carbon plus one hydrogen at each end of the chain:



A formula such as  $C_4H_{12}O_2$  for the unknown would be impossible, and the liquid (since it comes from wine) is ethyl alcohol,  $CH_3CH_2OH$ .

The Liebig method of analysis for carbon and hydrogen, and a similar method introduced by Jean Dumas<sup>6</sup> in 1830 for the analysis of nitrogen, were remarkable achievements at the time but were limited in their usefulness by the large sample sizes required. Often it is practically impossible to obtain more than milligram amounts of new compounds, and an analysis that destroys half-gram amounts of material at a time is unthinkable.

The major limiting factor in the Liebig and Dumas analyses was the accuracy of the analytical balances used to weigh the samples and the collection tubes. Under the leadership of Fritz Pregl,<sup>7</sup> however, a microbalance of great precision was developed, allowing highly accurate weighing of sub-milligram amounts. Pregl further refined all aspects of the Liebig method and, in 1911, introduced a method of microanalysis that could be carried out on 5–10 milligram (mg) samples. For his accomplishments, he received the 1923 Nobel Prize in chemistry.

Today, microanalysis of organic compounds is still carried out by the methods pioneered by Liebig, Dumas, and Pregl, although the techniques have become highly automated. Modern chemists do not consider a new compound to be fully characterized until accurate combustion analyses have been carried out, and many chemical journals still require such data before they publish new work.

PROBLEM .....

**2.20** Calculate the percentage of each element in these molecular formulas:

- |                                   |   |
|-----------------------------------|---|
| (a) Benzene, $C_6H_6$             | (b) Laetrile, $C_{14}H_{15}NO_7$          |
| (c) Quinine, $C_{20}H_{24}N_2O_2$ | (d) Diethylstilbestrol, $C_{18}H_{20}O_2$ |

<sup>6</sup>Jean Baptiste André Dumas (1800–1884); b. Alais, France; professor, École Polytechnique (1835), École de Médecine, Sorbonne.

<sup>7</sup>Fritz Pregl (1869–1930); b. Laibach, Austria; Ph.D. Graz (1893); professor, Innsbruck, Graz; Nobel Prize (1923).

## PROBLEM.....

- 2.21 The chemical responsible for the odor of lemon is a substance named citral. Combustion analysis shows citral to contain 78.9% C and 10.6% H. Assuming that the remainder is due to oxygen, what is the empirical formula of citral? If citral has a molecular weight of 152, what is its molecular formula?

## PROBLEM.....

- 2.22 A sample of squalene, isolated from shark oil, was submitted for combustion analysis. An amount weighing 8.00 mg gave 25.60 mg of CO<sub>2</sub> and 8.75 mg of H<sub>2</sub>O. Calculate the empirical formula of squalene. If squalene has a molecular weight of 410, what is its molecular formula?

## 2.10 Summary and Key Words

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Plus (+) and minus (−) signs are used to indicate the presence of **formal charges** on atoms. Assigning formal charges to specific atoms in neutral compounds is a bookkeeping technique that allows us to keep track of the valence electrons in an atom:

$$\text{Formal charge} = \left[ \begin{array}{c} \text{Number of valence electrons} \\ \text{in the free atom} \end{array} \right] - \left[ \begin{array}{c} \text{Number of valence electrons} \\ \text{in the bonded atom} \end{array} \right]$$

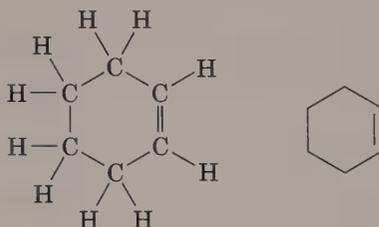
Organic molecules often have **polar covalent bonds** as a result of unsymmetrical electron sharing due to the intrinsic **electronegativity** of atoms. For example, a carbon–chlorine bond is polar because chlorine attracts the shared electrons more strongly than carbon does. Carbon–metal bonds, by contrast, are usually polarized in the opposite sense because carbon attracts electrons more strongly than most metals. Carbon–hydrogen bonds are relatively nonpolar. Many molecules as a whole are also polar owing to the cumulative effects of individual polar bonds, formal charges, and electron lone pairs. The polarity of a molecule is measured by its **dipole moment**,  $\mu$ .

Acidity and basicity are closely related to polarity and electronegativity. A **Brønsted–Lowry acid** is a compound that can donate a proton (hydrogen ion, H<sup>+</sup>), and a **Brønsted–Lowry base** is a compound that can accept a proton. The exact strength of a Brønsted–Lowry acid or base is expressed by its acidity constant,  $K_a$ . More useful is the Lewis definition of acids and bases. A **Lewis acid** is a compound that has a low-energy empty orbital that can accept an electron pair; BF<sub>3</sub>, AlCl<sub>3</sub>, and H<sup>+</sup> are examples. A **Lewis base** is a compound that donates an unshared electron pair; NH<sub>3</sub> and H<sub>2</sub>O are examples. Many organic molecules that contain oxygen and nitrogen are weak Lewis bases.

The analysis of organic compounds can be carried out accurately on milligram amounts of sample. The organic material is burned, and the

combustion products are weighed to give information that can be used to establish the empirical formula of an unknown.

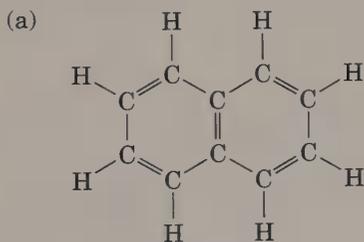
Chemists normally draw **skeletal structures** using a shorthand method in which carbons and most hydrogen atoms are not indicated. A carbon atom is assumed to be at the ends and at the junctions of lines (bonds), and the correct number of hydrogens is mentally supplied. For example:



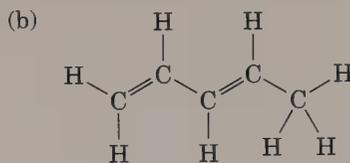
Cyclohexene

## ADDITIONAL PROBLEMS .....

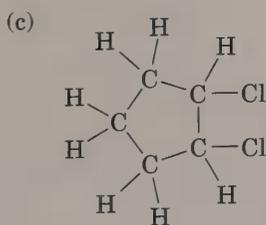
2.23 Convert the following structures into skeletal drawings:



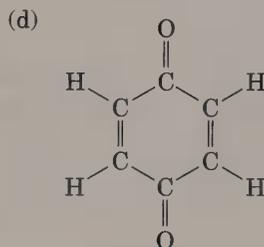
Naphthalene



1,3-Pentadiene

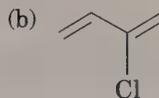
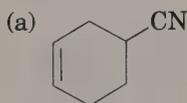


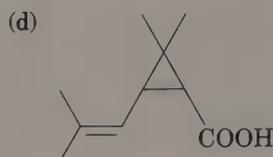
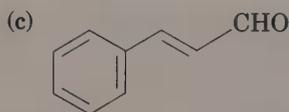
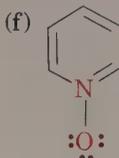
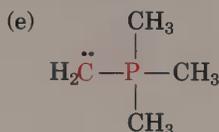
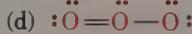
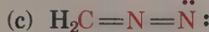
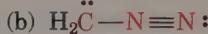
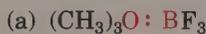
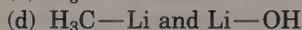
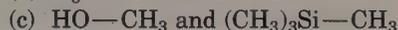
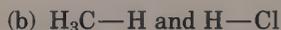
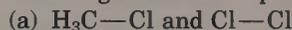
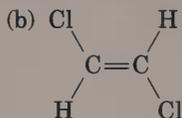
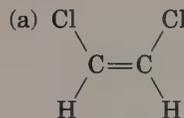
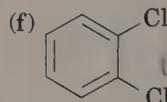
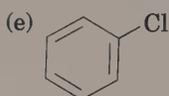
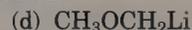
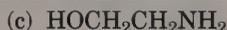
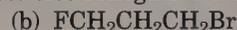
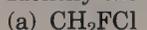
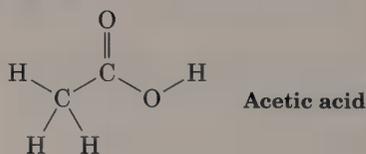
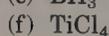
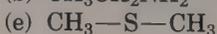
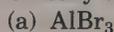
1,2-Dichlorocyclopentane



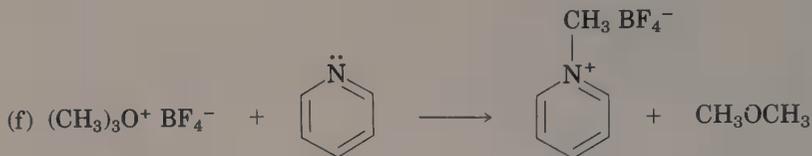
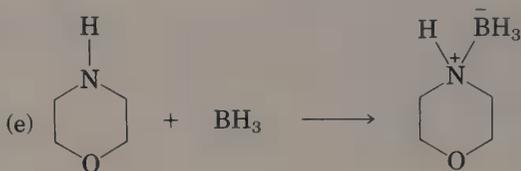
Quinone

2.24 Convert these skeletal drawings into Kekulé structures that show all carbons and hydrogens:



**2.25** Calculate the formal charges on the atoms indicated:**2.26** Use the electronegativity table (Figure 2.3) to predict which bond in each of the following sets is more polar:**2.27** Indicate the direction of bond polarity for each compound in Problem 2.26.**2.28** Which of these molecules have dipole moments? Indicate the expected direction of each.*cis*-1,2-Dichloroethylene*trans*-1,2-Dichloroethylene**2.29** How can you explain the fact that fluoromethane ( $\text{CH}_3\text{F}$ ,  $\mu = 1.81$  D) has a smaller dipole moment than chloromethane ( $\text{CH}_3\text{Cl}$ ,  $\mu = 1.87$  D) even though fluorine is more electronegative than chlorine?**2.30** Identify the most electronegative element in each of these molecules:**2.31** Alcohols can act either as weak acids or as weak bases, just as water can. Formulate the reactions of methyl alcohol,  $\text{CH}_3\text{OH}$ , with a strong acid such as  $\text{HCl}$  and with a strong base such as  $\text{Na}^+ \text{ } ^-\text{NH}_2$ .**2.32** How can you explain the fact that the O-H hydrogen in acetic acid is more acidic than any of the C-H hydrogens? [Hint: Consider bond polarity.]**2.33** Classify the following reagents as either Lewis acids or Lewis bases:

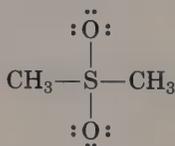




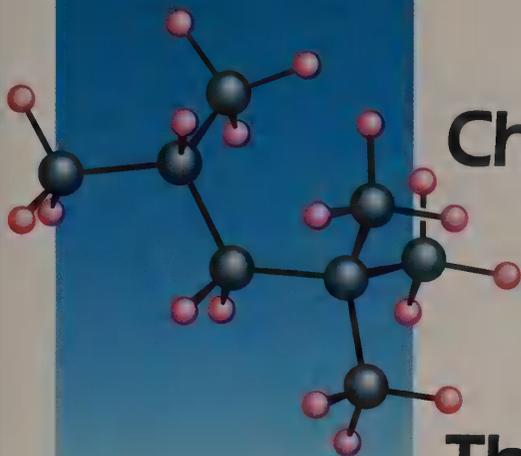
- 2.46** Calculate the percentage of each element in the following formulas:  
 (a) Aspirin,  $C_9H_8O_4$  (b) Muscone (musk oil),  $C_{16}H_{30}O$   
 (c) Morphine,  $C_{17}H_{19}NO_3$  (d) Strychnine,  $C_{21}H_{22}N_2O_2$
- 2.47**  $\alpha$ -Pinene, the main constituent of turpentine, has been shown by mass spectroscopy to have a molecular weight of 136. Combustion analytical data indicate that  $\alpha$ -pinene contains 88.3% C and 11.6% H by weight. What is the molecular formula of  $\alpha$ -pinene?
- 2.48** Jasmone, an odoriferous compound isolated from the jasmine flower, is valued for its use in perfumery. A pure sample of jasmone is analyzed and found to have 80.7% C and 9.7% H. Its molecular weight is 164. What other element is probably present in jasmone? What is the molecular formula of jasmone?
- 2.49** Progesterone, the so-called pregnancy hormone, was isolated by Adolf Butenandt in 1934. By extracting the ovaries of 50,000 sows, Butenandt was able to isolate 20 mg of the pure hormone. Microanalysis of a 0.00500 g sample by the Pregl techniques leads to the production of 0.0147 g  $CO_2$  and 0.0041 g  $H_2O$ . What is the empirical formula for progesterone? Since we now know the molecular weight of progesterone to be 314, what is the molecular formula?
- 2.50** The Dumas method of analysis for nitrogen content in a molecule involves measuring the amount of nitrogen gas that is produced in a complex degradation reaction. Application of the gas laws tells us that 28.0 mg  $N_2$  (1.00 millimole, mmol) has a volume of 22.4 milliliters (mL). By measuring the amount of  $N_2$  derived from a sample of known weight, we can arrive at a value for the percentage of nitrogen in the sample.

Cadaverine, an aptly named amine with a molecular weight of 102, was analyzed for C, H, and N. Analysis of a 0.040 g sample yielded 0.086 g  $CO_2$ , 0.051 g  $H_2O$ , and 8.6 mL  $N_2$  gas collected at standard temperature and pressure. How many grams of  $N_2$  gas were produced? What is the percentage of N in the 0.040 g sample? What is the molecular formula for cadaverine? What do you think cadaverine smells like?

- 2.51** Dimethyl sulfone has a high dipole moment ( $\mu = 4.4$  D). Calculate the formal charges present on oxygen and sulfur, and suggest a geometry that is consistent with the observed dipole moment.



Dimethyl sulfone



## Chapter 3

# The Nature of Organic Compounds: Alkanes and Cycloalkanes

According to *Chemical Abstracts*, the invaluable publication that abstracts and indexes the chemical literature, there are more than 10 million known organic compounds. Each of these compounds has its own physical properties such as melting point and boiling point, and each has its own chemical reactivity.

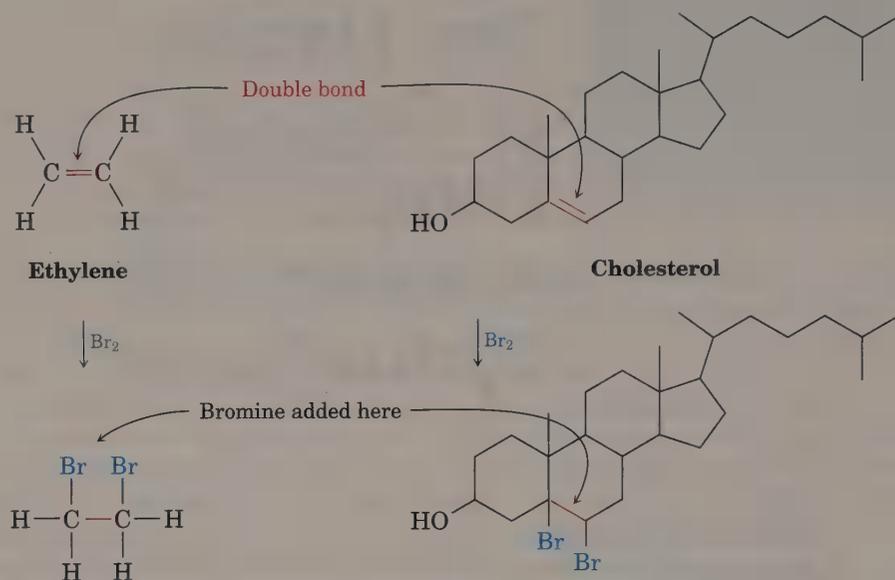
Chemists have learned through many years of experience that organic compounds can be classified into families according to their structural features and that the chemical reactivity of the members of a given family is often predictable. Instead of 10 million compounds with random reactivity, there are several dozen general families of organic compounds whose chemistry is roughly predictable. We'll study the chemistry of the most important families throughout the rest of this book, beginning in the present chapter with a look at the simplest one, the *alkanes*.

### 3.1 Functional Groups

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The structural features that make it possible to classify compounds by reactivity are called functional groups. A **functional group** is a part of a larger molecule and is composed of an atom or a group of atoms that have a characteristic chemical behavior. Chemically, a given functional group behaves approximately the same way in every molecule it's a part of. For example, one of the simplest functional groups is the carbon–carbon double

bond. Since the electronic structure of the carbon–carbon double bond remains essentially the same in all molecules where it occurs, its chemical reactivity also remains the same. Ethylene, the simplest compound with a double bond, undergoes reactions that are remarkably similar to those of cholesterol, a much more complicated molecule. Both, for example, react with bromine to give products in which a bromine atom has added to each of the double-bond carbons (Figure 3.1).



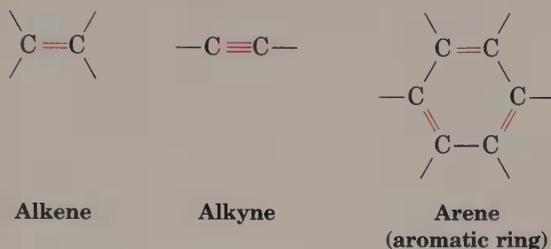
**Figure 3.1** The reactions of ethylene and cholesterol with bromine. In both cases, bromine reacts with the C=C double-bond functional group in exactly the same way. The size and nature of the remainder of the molecule are unimportant.

The example shown in Figure 3.1 is typical: *The chemistry of every organic molecule, regardless of size and complexity, is determined by the functional groups it contains.* Table 3.1 lists many of the common functional groups and gives simple examples of their occurrence. Look carefully at this table to see the many types of functional groups found in organic compounds. Some functional groups, such as those in alkenes, alkynes, and aromatic rings, have only carbon–carbon double or triple bonds; others have halogen; and still others have oxygen, nitrogen, or sulfur. Much of the chemistry you'll be studying in the remainder of this book is the chemistry of these functional groups.

It's a good idea at this point to memorize the structures of the functional groups shown in Table 3.1 so that they'll be familiar when you see them again. They can be grouped into several categories.

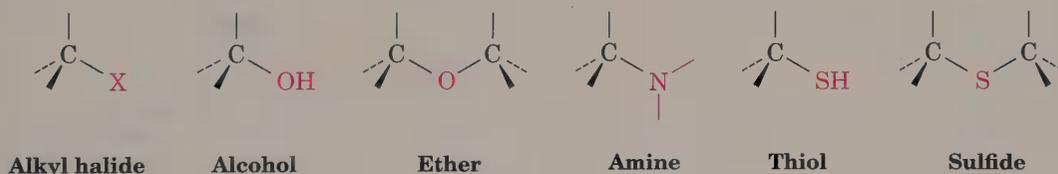
## FUNCTIONAL GROUPS WITH CARBON–CARBON MULTIPLE BONDS

Alkenes, alkynes, and arenes (aromatic compounds) all contain carbon–carbon multiple bonds. Alkenes have a double bond, alkynes have a triple bond, and aromatic rings have three alternating double and single bonds in a six-membered ring of carbon atoms. Because of their structural similarities, these compounds also have chemical similarities.



## FUNCTIONAL GROUPS WITH CARBON SINGLY BONDED TO AN ELECTRONEGATIVE ATOM

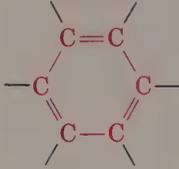
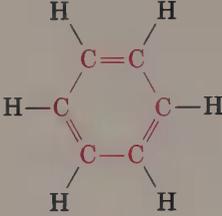
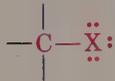
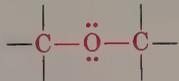
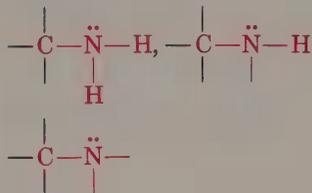
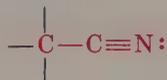
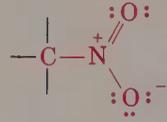
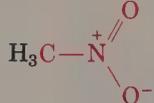
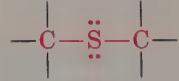
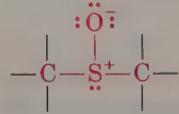
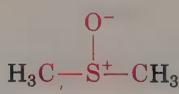
Alkyl halides, alcohols, ethers, amines, sulfides, thiols, and several others all have a carbon atom singly bonded to an electronegative atom—a halogen, an oxygen, a nitrogen, or a sulfur. Alkyl halides have a carbon atom bonded to halogen, alcohols have a carbon atom bonded to a hydroxyl (–OH) group, ethers have two carbon atoms bonded to the same oxygen, amines have a carbon atom bonded to a nitrogen, thiols have a carbon atom bonded to an –SH group, and sulfides have two carbon atoms bonded to the same sulfur. In all cases, the bonds are polar with the carbon atom bearing a slight positive charge ( $\delta^+$ ) and the electronegative atom bearing a slight negative charge ( $\delta^-$ ).



## FUNCTIONAL GROUPS WITH A CARBON–OXYGEN DOUBLE BOND (CARBONYL GROUPS)

Note particularly in Table 3.1 the different families of compounds that contain the **carbonyl group**, C=O (pronounced car-bo-neel). Carbon–oxygen double bonds are present in some of the most important compounds in organic chemistry. These compounds are similar in many respects but differ depending on the identity of the atoms bonded to the carbonyl-group carbon. Aldehydes have one carbon and one hydrogen bonded to the C=O, ketones have two carbons bonded to the C=O, carboxylic acids have one carbon and

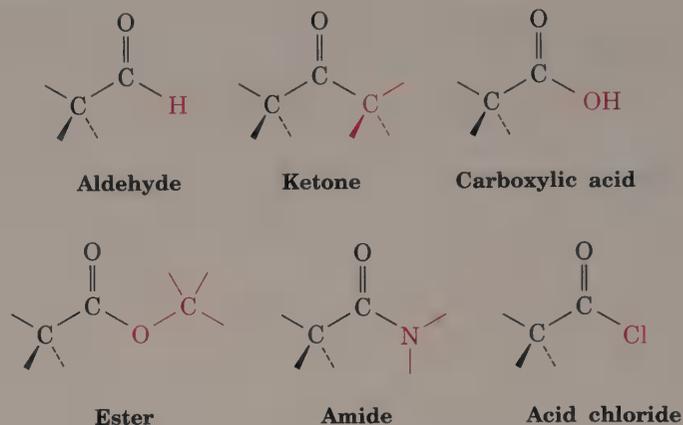
**TABLE 3.1** Structures of Some Important Functional Groups

<i>Family name</i>	<i>Functional group structure<sup>a</sup></i>	<i>Simple example</i>	<i>Name ending</i>
Alkane	(Contains only C—H and C—C single bonds)	CH <sub>3</sub> CH <sub>3</sub>	-ane Ethane
Alkene		H <sub>2</sub> C=CH <sub>2</sub>	-ene Ethene (Ethylene)
Alkyne	—C≡C—	H—C≡C—H	-yne Ethyne (Acetylene)
Arene			None Benzene
Halide	 (X = F, Cl, Br, I)	H <sub>3</sub> C—Cl	None Chloromethane
Alcohol		H <sub>3</sub> C—O—H	-ol Methanol
Ether		H <sub>3</sub> C—O—CH <sub>3</sub>	ether Dimethyl ether
Amine		H <sub>3</sub> C—NH <sub>2</sub>	-amine Methylamine
Nitrile		H <sub>3</sub> C—C≡N	-nitrile Ethanenitrile (Acetonitrile)
Nitro			None Nitromethane
Sulfide		H <sub>3</sub> C—S—CH <sub>3</sub>	sulfide Dimethyl sulfide
Sulfoxide			sulfoxide Dimethyl sulfoxide

<sup>a</sup>The bonds whose connections aren't specified are assumed to be attached to carbon or hydrogen atoms in the rest of the molecule.

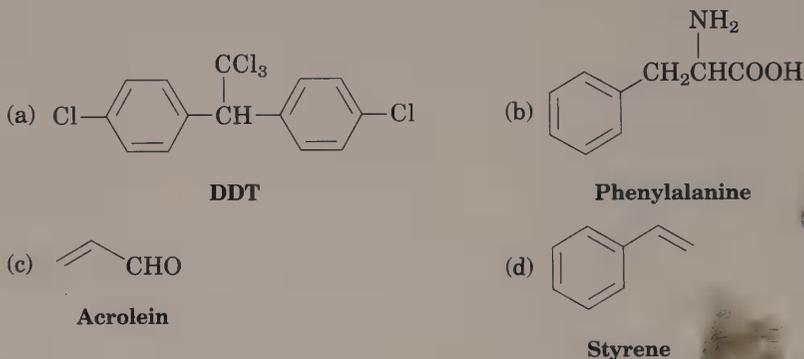
Family name	Functional group structure <sup>a</sup>	Simple example	Name ending
Sulfone	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\   \\ -\text{C}-\text{S}^{2+}-\text{C}- \\   \\ \text{:}\ddot{\text{O}}\text{:} \end{array}$	$\begin{array}{c} \text{O}^- \\   \\ \text{H}_3\text{C}-\text{S}^{2+}-\text{CH}_3 \\   \\ \text{O}^- \end{array}$	<i>sulfone</i> Dimethyl sulfone
Thiol	$\begin{array}{c}   \\ -\text{C}-\ddot{\text{S}}-\text{H} \\   \end{array}$	$\text{H}_3\text{C}-\text{SH}$	<i>-thiol</i> Methanethiol
Carbonyl,	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}- \end{array}$		
Aldehyde	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\text{H} \\   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{H} \end{array}$	<i>-al</i> Ethanal (Acetaldehyde)
Ketone	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\text{C}- \\   \quad   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{CH}_3 \end{array}$	<i>-one</i> Propanone (Acetone)
Carboxylic acid	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\ddot{\text{O}}\text{H} \\   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{OH} \end{array}$	<i>-oic acid</i> Ethanoic acid (Acetic acid)
Ester	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\ddot{\text{O}}-\text{C}- \\   \quad   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{O}-\text{CH}_3 \end{array}$	<i>-oate</i> Methyl ethanoate (Methyl acetate)
Amide	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\ddot{\text{N}}\text{H}_2 \\   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{NH}_2 \end{array}$	<i>-amide</i> Ethanamide (Acetamide)
	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\ddot{\text{N}}-\text{H} \\   \end{array}$		
	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\ddot{\text{N}}- \\   \end{array}$		
Carboxylic acid chloride	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \\    \\ -\text{C}-\text{C}-\text{Cl} \\   \end{array}$	$\begin{array}{c} \text{O} \\    \\ \text{H}_3\text{C}-\text{C}-\text{Cl} \end{array}$	<i>-oyl chloride</i> Ethanoyl chloride (Acetyl chloride)
Carboxylic acid anhydride	$\begin{array}{c} \text{:}\ddot{\text{O}}\text{:} \quad \text{:}\ddot{\text{O}}\text{:} \\    \quad    \\ -\text{C}-\text{C}-\ddot{\text{O}}-\text{C}-\text{C}- \\   \quad   \end{array}$	$\begin{array}{c} \text{O} \quad \text{O} \\    \quad    \\ \text{H}_3\text{C}-\text{C}-\text{O}-\text{C}-\text{CH}_3 \end{array}$	<i>-oic anhydride</i> Ethanoic anhydride (Acetic anhydride)

one  $\text{-OH}$  group bonded to the  $\text{C=O}$ , esters have one carbon and one ether-like oxygen bonded to the  $\text{C=O}$ , amides have one carbon and one amine-like nitrogen bonded to the  $\text{C=O}$ , acid chlorides have a chlorine bonded to the  $\text{C=O}$ , and so on.



**PROBLEM** .....

3.1 Circle and identify the functional groups in each of the following molecules:



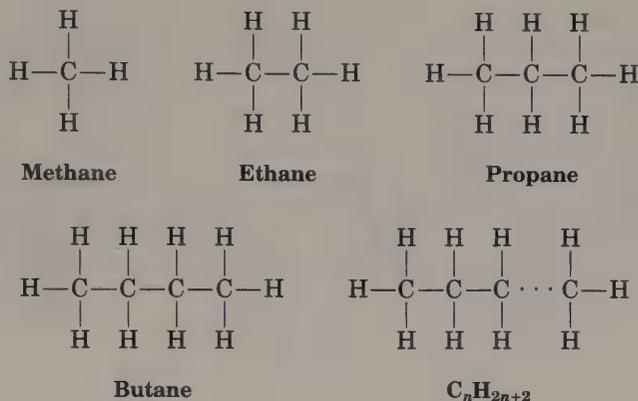
**PROBLEM** .....

3.2 Propose structures for simple molecules that contain these functional groups:

- |             |                           |                      |
|-------------|---------------------------|----------------------|
| (a) Alcohol | (b) Aromatic ring         | (c) Carboxylic acid  |
| (d) Amine   | (e) Both ketone and amine | (f) Two double bonds |
- .....

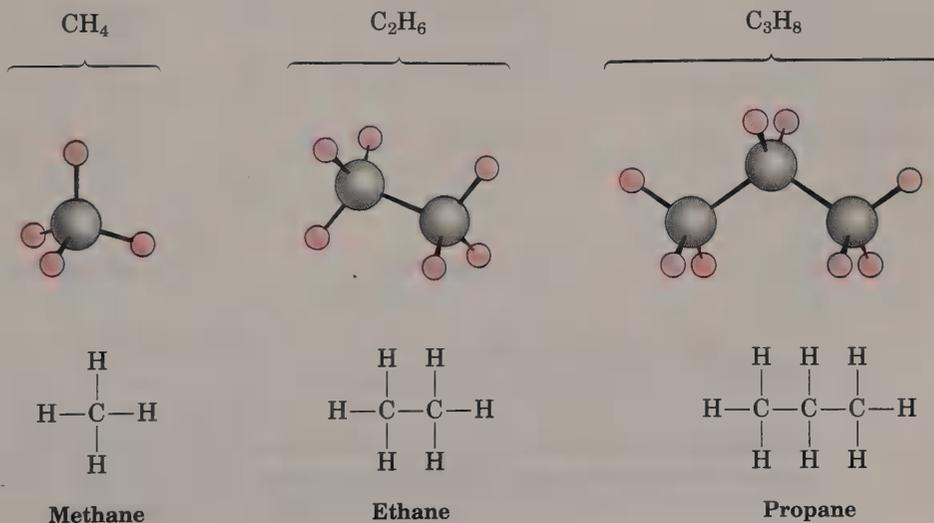
## 3.2 Alkanes and Alkane Isomers

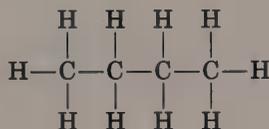
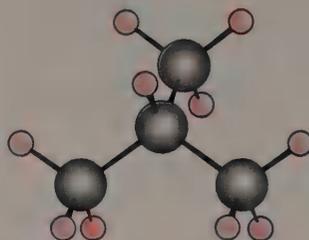
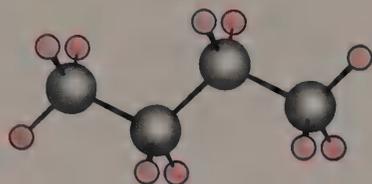
We saw in Section 1.9 that the carbon–carbon single bond in ethane results from overlap of two carbon  $sp^3$  orbitals. If we imagine joining three, four, five, or even more carbon atoms together by carbon–carbon single bonds, we can generate the large family of molecules called **alkanes**:



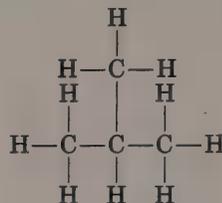
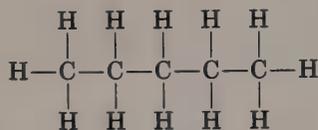
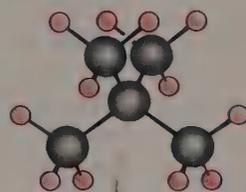
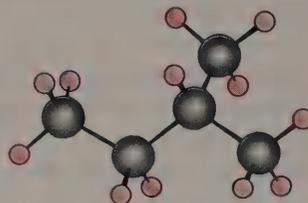
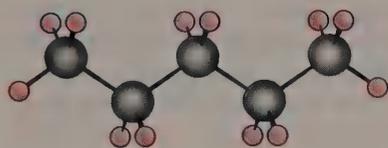
Alkanes are often described as saturated hydrocarbons—**hydrocarbons** because they contain only carbon and hydrogen; **saturated** because they have only C–C and C–H single bonds and thus contain the maximum possible number of hydrogens per carbon. They have the general formula  $\text{C}_n\text{H}_{2n+2}$ , where  $n$  is any integer. Alkanes are also occasionally referred to as **aliphatic** compounds, a name derived from the Greek *aleiphas*, meaning “fat.” We’ll see later that animal fats do indeed contain long carbon chains similar to alkanes.

Think for a moment about the ways that carbon and hydrogen can combine to make alkanes. With one carbon and four hydrogens, only one structure is possible: methane,  $\text{CH}_4$ . Similarly, there is only one possible combination of two carbons with six hydrogens (ethane,  $\text{CH}_3\text{CH}_3$ ) and only one possible combination of three carbons with eight hydrogens (propane,  $\text{CH}_3\text{CH}_2\text{CH}_3$ ). If larger numbers of carbons and hydrogens combine, however, more than one kind of molecule can form. For example, there are two ways that molecules with the formula  $\text{C}_4\text{H}_{10}$  can form: The four carbons can be in a row (butane), or they can branch (isobutane). Similarly, there are three ways in which  $\text{C}_5\text{H}_{12}$  molecules can form, and so on for larger alkanes:

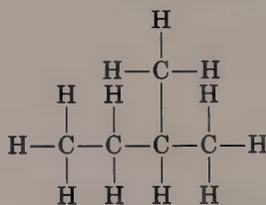




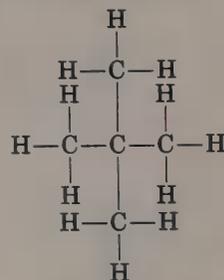
Butane

Isobutane  
(2-methylpropane)

Pentane



2-Methylbutane



2,2-Dimethylpropane

Compounds like butane, whose carbons are connected in a row, are called **straight-chain alkanes**, or **normal alkanes**, whereas compounds whose carbon chains branch, such as 2-methylpropane (isobutane), are called **branched-chain alkanes**. Note that you can draw a line connecting all the carbons of a straight-chain alkane without retracing your path or lifting your pencil from the paper. For a branched-chain alkane, however, you have

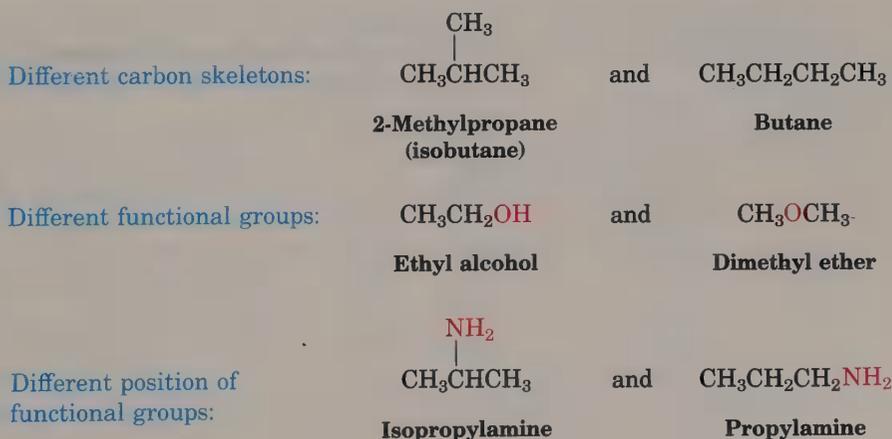
to either retrace your path or lift your pencil from the paper in order to draw a line connecting all the carbons.

Compounds like the two  $C_4H_{10}$  molecules that have the same formula but different structures are called *isomers*, from Greek *isos* + *meros* meaning “made of the same parts.” **Isomers** are compounds that have the same numbers and kinds of atoms but that differ in the way these atoms are arranged. Compounds like butane and isobutane, whose atoms are connected differently, are called **constitutional isomers**. We’ll see shortly that other kinds of isomerism are also possible, even among compounds whose atoms are connected in the same order. As Table 3.2 shows, the number of possible alkane isomers increases dramatically as the number of carbon atoms increases.

**TABLE 3.2** Number of Alkane Isomers

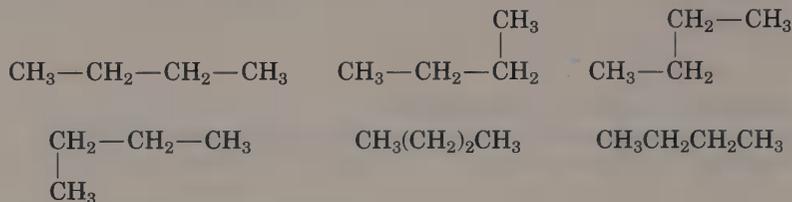
<i>Formula</i>	<i>Number of isomers</i>	<i>Formula</i>	<i>Number of isomers</i>
$C_6H_{14}$	5	$C_{10}H_{22}$	75
$C_7H_{16}$	9	$C_{15}H_{32}$	4,347
$C_8H_{18}$	18	$C_{20}H_{42}$	366,319
$C_9H_{20}$	35	$C_{30}H_{62}$	4,111,846,763

Constitutional isomerism is not limited to alkanes—it occurs widely throughout organic chemistry. Constitutional isomers may have different carbon skeletons (as in isobutane and butane), different functional groups (as in ethyl alcohol and dimethyl ether), or different locations of a functional group along the chain (as in isopropylamine and propylamine). Regardless of the reason for the isomerism, constitutional isomers always have the same formula but a different connection of atoms:



A given alkane can be arbitrarily drawn in many ways. For example, the straight-chain, four-carbon alkane called butane can be represented by any of the structures shown in Figure 3.2. These structures don’t imply any

particular three-dimensional geometry for butane; they only indicate the connections among its atoms. In practice, we soon tire of drawing all the bonds in a molecule and usually refer to butane by the shorthand **condensed structure**,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ , or even more simply as  $n\text{-C}_4\text{H}_{10}$ , where  $n$  signifies *normal*, straight-chain butane. In condensed structures like  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$ , the C–C and C–H bonds aren't usually shown but are "understood." If a carbon has three hydrogens bonded to it, we write  $\text{CH}_3$ ; if a carbon has two hydrogens bonded to it, we write  $\text{CH}_2$ , and so on.



**Figure 3.2** Some representations of butane,  $\text{C}_4\text{H}_{10}$ . The molecule is the same regardless of how it's drawn. These structures imply only that butane has a continuous chain of four carbon atoms.

Straight-chain alkanes are named according to the number of carbon atoms in the chain, as shown in Table 3.3. With the exception of the first four compounds—methane, ethane, propane, and butane—whose names have historical roots, the alkanes are named using Greek numbers. The suffix *-ane* is added to the end of each name to indicate that the molecule identified is an alkane. Thus, *pentane* is the five-carbon alkane, *hexane* is the six-carbon alkane, and so on. The names of at least the first ten should be memorized.

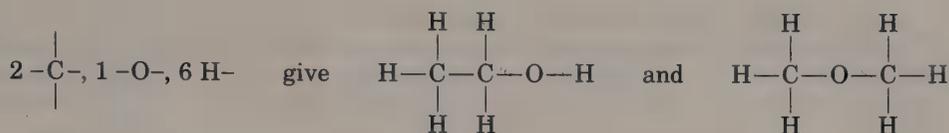
**TABLE 3.3** Names of Straight-Chain Alkanes

Number of carbons ( $n$ )	Name	Formula ( $\text{C}_n\text{H}_{2n+2}$ )	Number of carbons ( $n$ )	Name	Formula ( $\text{C}_n\text{H}_{2n+2}$ )
1	Methane	$\text{CH}_4$	9	Nonane	$\text{C}_9\text{H}_{20}$
2	Ethane	$\text{C}_2\text{H}_6$	10	Decane	$\text{C}_{10}\text{H}_{22}$
3	Propane	$\text{C}_3\text{H}_8$	11	Undecane	$\text{C}_{11}\text{H}_{24}$
4	Butane	$\text{C}_4\text{H}_{10}$	12	Dodecane	$\text{C}_{12}\text{H}_{26}$
5	Pentane	$\text{C}_5\text{H}_{12}$	13	Tridecane	$\text{C}_{13}\text{H}_{28}$
6	Hexane	$\text{C}_6\text{H}_{14}$	20	Icosane	$\text{C}_{20}\text{H}_{42}$
7	Heptane	$\text{C}_7\text{H}_{16}$	21	Henicosane	$\text{C}_{21}\text{H}_{44}$
8	Octane	$\text{C}_8\text{H}_{18}$	30	Triacotane	$\text{C}_{30}\text{H}_{62}$

**PRACTICE PROBLEM**.....

Propose structures for two isomers of formula  $\text{C}_2\text{H}_6\text{O}$ .

**Solution** We know that carbon forms four bonds, oxygen forms two, and hydrogen forms one. Putting the pieces together yields two isomeric structures:



PROBLEM .....

3.3 Draw structures of the five isomers of  $\text{C}_6\text{H}_{14}$ .

PROBLEM .....

3.4 There are seven constitutional isomers with the formula  $\text{C}_4\text{H}_{10}\text{O}$ . Draw as many as you can.

PROBLEM .....

3.5 Propose structures that meet the following descriptions:

(a) Two isomeric esters with formula  $\text{C}_5\text{H}_{10}\text{O}_2$

(b) Two isomeric nitriles with formula  $\text{C}_4\text{H}_7\text{N}$

PROBLEM .....

3.6 How many isomers are there that have the following structures?

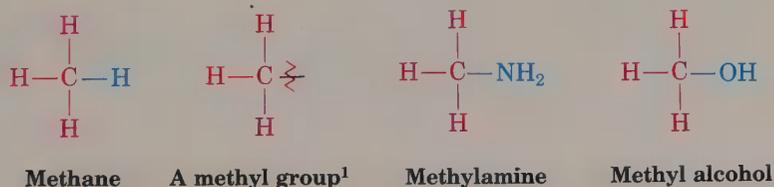
(a) Alcohols with formula  $\text{C}_3\text{H}_8\text{O}$

(b) Bromoalkanes with formula  $\text{C}_4\text{H}_9\text{Br}$

.....

### 3.3 Alkyl Groups

If a hydrogen atom is removed from an alkane, the part-structure that remains is called an **alkyl group**. Alkyl groups are named by replacing the *-ane* ending of the parent alkane with an *-yl* ending. For example, removal of a hydrogen from methane,  $\text{CH}_4$ , generates a *methyl* group,  $-\text{CH}_3$ , and removal of a hydrogen from ethane,  $\text{CH}_3\text{CH}_3$ , generates an *ethyl* group,  $-\text{CH}_2\text{CH}_3$ . Similarly, removal of a hydrogen atom from the end carbon of any *n*-alkane gives the series of straight-chain *n*-alkyl groups shown in Table 3.4. The combination of an alkyl group with any of the functional groups listed earlier allows us to generate and name many hundreds of thousands of compounds.



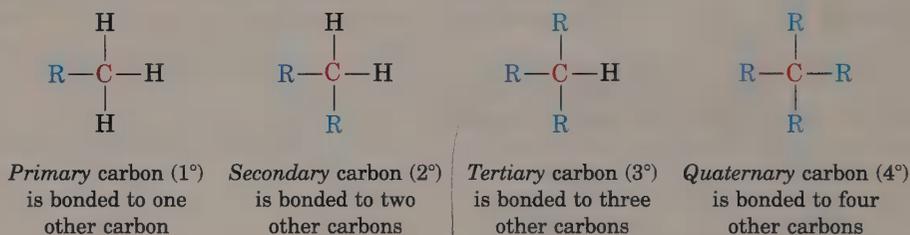
<sup>1</sup>The symbol  $\text{X}$  will be used throughout this book to indicate that the partial organic structure shown is bonded to another, unspecified group.

TABLE 3.4 Some Straight-Chain Alkyl Groups

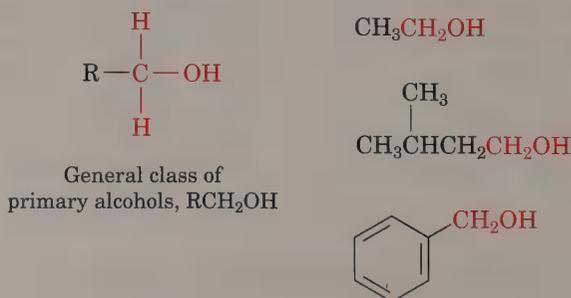
Alkane	Name	Alkyl group	Name (abbreviation)
CH <sub>4</sub>	Methane	—CH <sub>3</sub>	Methyl (Me)
CH <sub>3</sub> CH <sub>3</sub>	Ethane	—CH <sub>2</sub> CH <sub>3</sub>	Ethyl (Et)
CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub>	Propane	—CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	Propyl (Pr)
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	Butane	—CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	Butyl (Bu)
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	Pentane	—CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	Pentyl

Just as *n*-alkyl groups are generated by removing a hydrogen from an *end* carbon, branched alkyl groups are generated by removing an *internal* hydrogen atom. Two 3-carbon alkyl groups and four 4-carbon alkyl groups are possible (Figure 3.3).

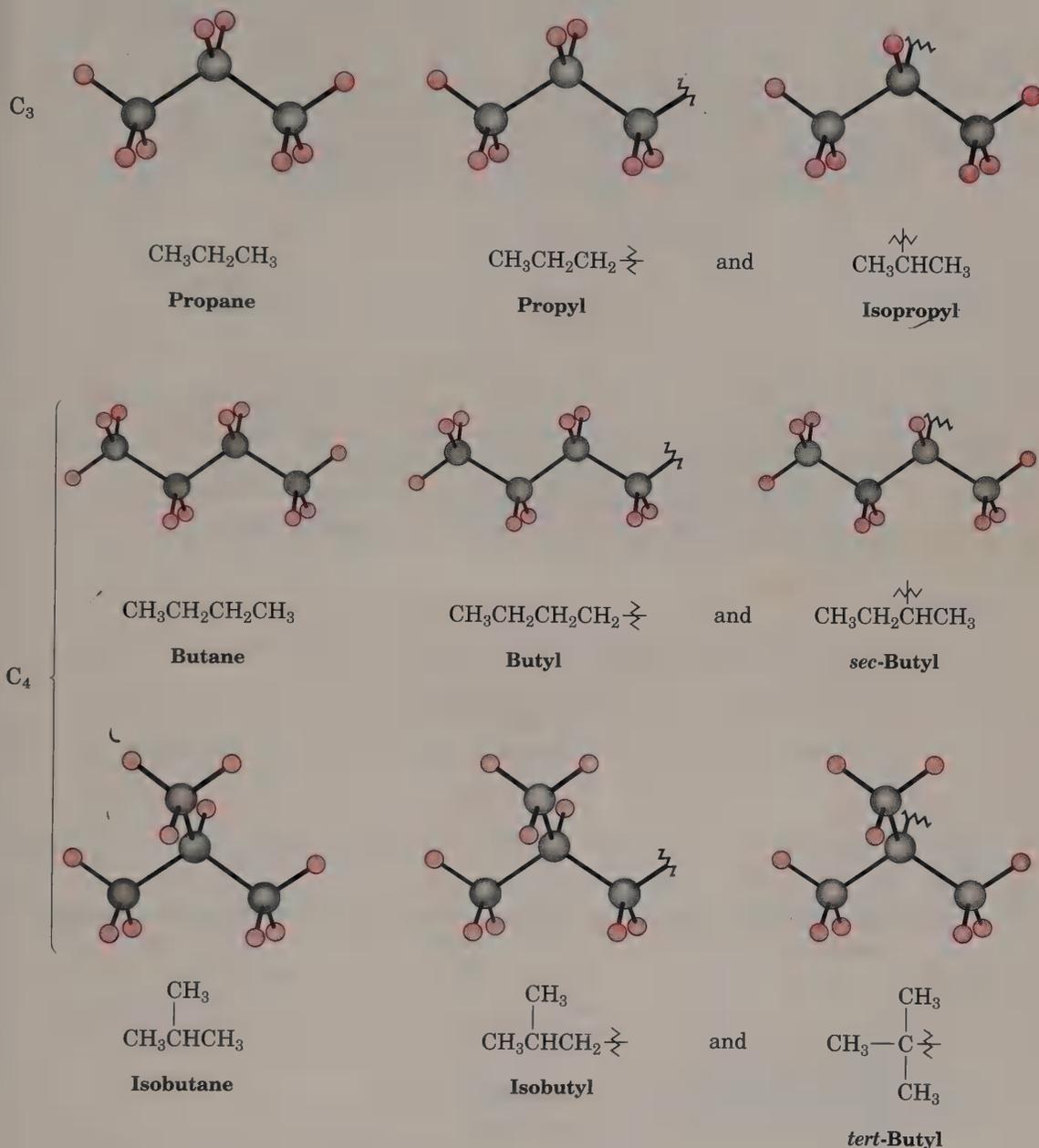
One word of explanation about naming alkyl groups: The prefixes used for the C<sub>4</sub> alkyl groups in Figure 3.3, *sec* (for secondary) and *tert* (for tertiary), refer to the degree of alkyl substitution at the carbon atom in question. There are four possible degrees of alkyl substitution for carbon, denoted 1° (primary), 2° (secondary), 3° (tertiary), and 4° (quaternary):



The symbol **R** is used to represent a generalized alkyl group. The **R** group can be methyl, ethyl, propyl, or any of an infinite number of others. You might think of **R** as representing the **R**est of the molecule, which we aren't bothering to specify because it's not important. The terms *primary*, *secondary*, *tertiary*, and *quaternary* are routinely used in organic chemistry, and their meanings must become second nature. For example, if we were to say "The product of the reaction is a primary alcohol," we would be talking about the *general class of compounds* that has an alcohol functional group (—OH) bonded to a primary carbon atom, RCH<sub>2</sub>OH:

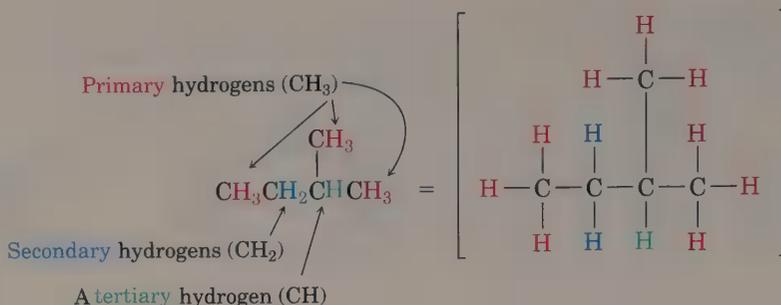


Some specific examples of  
primary alcohols, RCH<sub>2</sub>OH



**Figure 3.3** Generation of straight-chain and branched-chain alkyl groups from *n*-alkanes.

In addition, we also speak sometimes about hydrogen atoms as being primary, secondary, or tertiary. Primary hydrogen atoms are on  $CH_3$  groups, secondary hydrogens are on  $CH_2$  groups, and tertiary hydrogens are on  $CH$  groups:



PROBLEM.....

3.7 Draw the eight five-carbon alkyl groups (pentyl isomers).

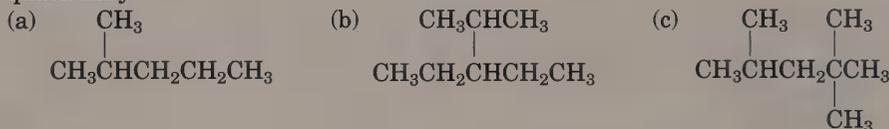
PROBLEM.....

3.8 Draw alkanes that meet these descriptions:

- An alkane with two tertiary carbons
- An alkane that contains an isopropyl group
- An alkane that has one quaternary and one secondary carbon

PROBLEM.....

3.9 Identify the carbon atoms in these molecules as primary, secondary, tertiary, or quaternary:



PROBLEM.....

3.10 Identify the hydrogen atoms on the compounds shown in Problem 3.9 as primary, secondary, or tertiary.

.....

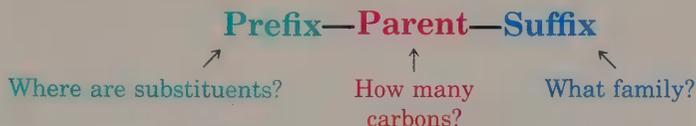
## 3.4 Naming Alkanes

In earlier times, when relatively few pure organic chemicals were known, new compounds were named at the whim of their discoverer. Thus, urea ( $\text{CH}_4\text{N}_2\text{O}$ ) is a pure crystalline substance isolated from urine; morphine ( $\text{C}_{17}\text{H}_{19}\text{NO}_3$ ) is an analgesic (painkiller) isolated in 1805 from the opium poppy and named after Morpheus, the Greek god of dreams; and barbituric acid is a tranquilizing agent named by its discoverer in honor of his friend Barbara.

As the science of organic chemistry slowly grew in the nineteenth century, so too did the need for a systematic method of unambiguously naming organic compounds. The system of nomenclature (naming) we'll use in this

book is that devised by the International Union of Pure and Applied Chemistry (IUPAC, usually spoken as **eye-you-pac**). IUPAC rules can deal with all functional groups and can name all but the most complex structures.

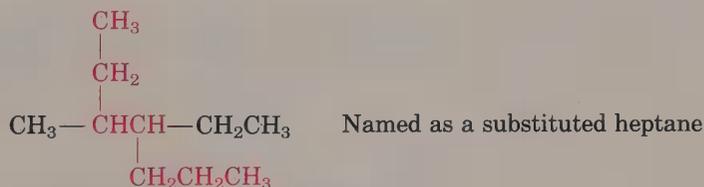
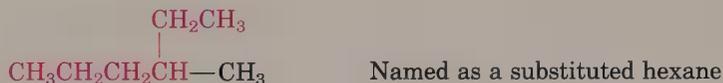
A chemical name has three parts in the IUPAC system: prefix, parent, and suffix. The parent tells how many carbon atoms are in the main chain and indicates the overall length of that chain, the suffix identifies the functional groups present in the molecule, and the prefix specifies the location of the functional groups and other substituents on the main chain:



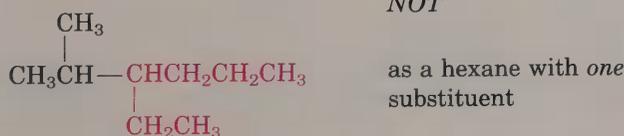
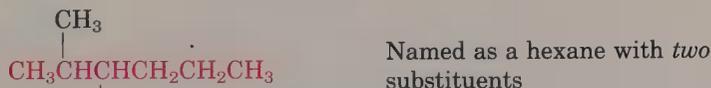
As we cover new functional groups in later chapters, the applicable IUPAC rules of nomenclature will be given. In addition, Appendix A gives an overall view of organic nomenclature and shows how compounds that contain more than one functional group can be named. For the present, though, let's see how we can name branched-chain alkanes.

All but the most complex branched-chain alkanes can be named by following four steps. For a very few compounds, a fifth step is needed.

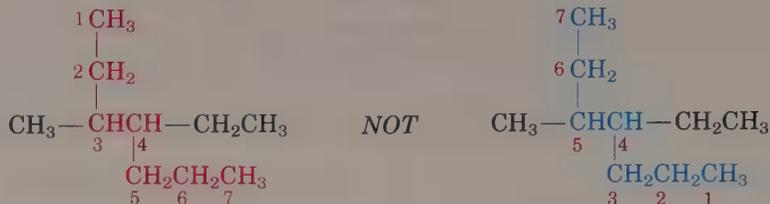
1. Find the parent hydrocarbon:
  - a. Find the *longest continuous carbon chain* present in the molecule, and use the name of that chain as the parent name. The longest chain may not always be apparent from the manner of writing; you may have to “turn corners.”



- b. If two different chains of equal length are present, choose the one with the larger number of branch points as the parent:



2. Number the atoms in the main chain.
- a. Beginning at the end *nearer the first branch point*, number each carbon atom in the longest chain you identified:



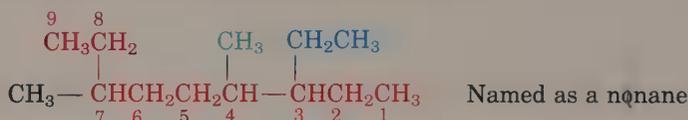
The first branch occurs at C3 in the proper numbering system but at C4 in the improper system.

- b. If there is branching an equal distance away from both ends of the parent chain, begin numbering at the end nearer the *second* branch point:



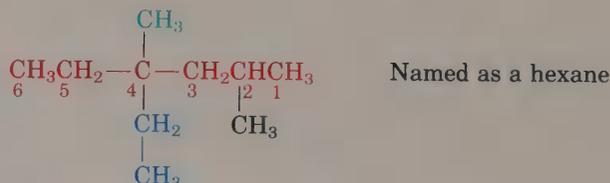
3. Identify and number the substituents.

- a. Using the numbering system you've decided is correct, assign a number to each substituent according to its point of attachment to the main chain:



Substituents: On C3, CH<sub>2</sub>CH<sub>3</sub> (3-ethyl)  
 On C4, CH<sub>3</sub> (4-methyl)  
 On C7, CH<sub>3</sub> (7-methyl)

- b. If there are two substituents on the same carbon, assign them both the same number. There must always be as many numbers in the name as there are substituents:

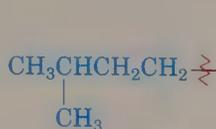


Substituents: On C2, CH<sub>3</sub> (2-methyl)  
 On C4, CH<sub>3</sub> (4-methyl)  
 On C4, CH<sub>2</sub>CH<sub>3</sub> (4-ethyl)

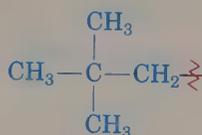




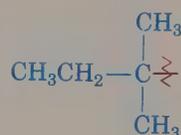
## 3. Five-carbon alkyl groups:



Isopentyl,  
also called  
isoamyl (*i*-amyl)

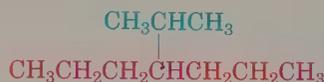


Neopentyl



*tert*-Pentyl,  
also called  
*tert*-amyl (*t*-amyl)

The common names of these simple alkyl groups are so well entrenched in the chemical literature that IUPAC rules make allowance for them. Thus, the following compound is properly named *either* 4-(1-methylethyl)heptane or 4-isopropylheptane. There's no choice but to memorize these common names; fortunately there aren't many of them.

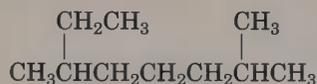


4-(1-Methylethyl)heptane or 4-Isopropylheptane

When writing an alkane name, the prefix *iso-* is considered part of the alkyl-group name for alphabetizing purposes, but the hyphenated prefixes *sec-* and *tert-* are not. Thus, isopropyl and isobutyl are listed alphabetically under *i*, but *sec*-butyl and *tert*-butyl are listed under *b*.

## PRACTICE PROBLEM.....

What is the IUPAC name of this alkane?

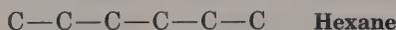


**Solution** The molecule has a chain of eight carbons (octane) with two methyl substituents. Numbering from the end nearer the first methyl substituent indicates that the methyls are at C2 and C6, giving the name 2,6-dimethyloctane.

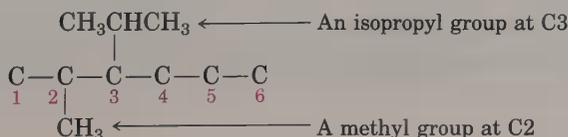
## PRACTICE PROBLEM.....

Draw the structure of 3-isopropyl-2-methylhexane.

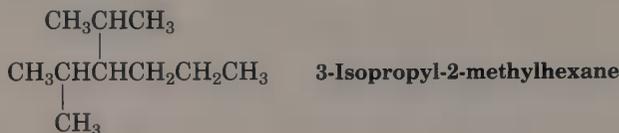
**Solution** First, look at the parent name (hexane) and draw its carbon structure:



Next, find the substituents (3-isopropyl and 2-methyl), and place them on the proper carbons:

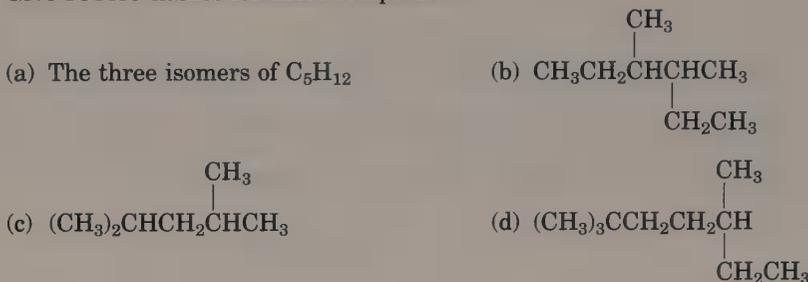


Finally, add hydrogens to complete the structure:



PROBLEM .....

3.11 Give IUPAC names for these compounds:



PROBLEM .....

3.12 Draw structures corresponding to these IUPAC names:

- |                                 |                                 |
|---------------------------------|---------------------------------|
| (a) 3,4-Dimethylnonane          | (b) 3-Ethyl-4,4-dimethylheptane |
| (c) 2,2-Dimethyl-4-propyloctane | (d) 2,2,4-Trimethylpentane      |

PROBLEM .....

3.13 The following names are incorrect. Draw the structures they represent, explain why the names are incorrect, and give correct names.

- |                                 |                             |
|---------------------------------|-----------------------------|
| (a) 1,1-Dimethylpentane         | (b) 3-Methyl-2-propylhexane |
| (c) 4,4-Dimethyl-3-ethylpentane | (d) 5-Ethyl-4-methylhexane  |
| (e) 2,3-Methylhexane            | (f) 3-Dimethylpentane       |

PROBLEM .....

3.14 Assign systematic names to the eight five-carbon alkyl groups (pentyl isomers).

.....

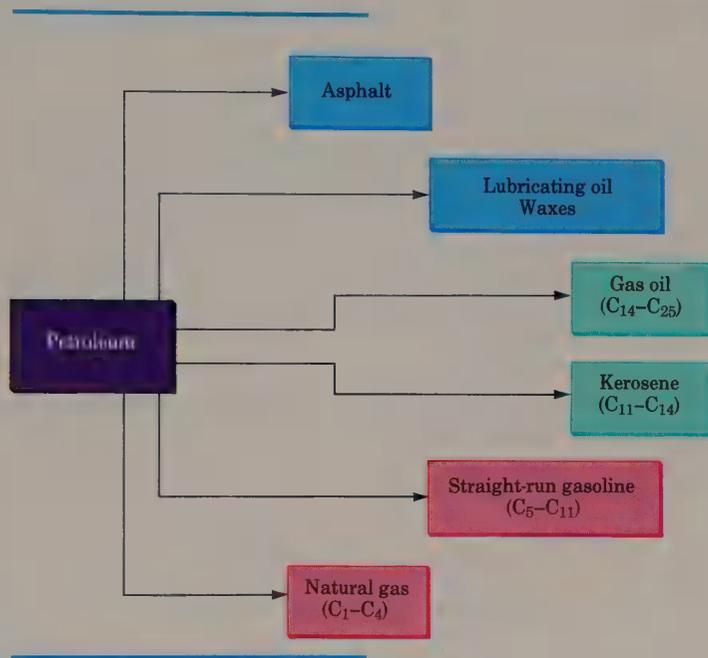
### 3.5 Occurrence of Alkanes: Petroleum

Many alkanes occur naturally in the plant and animal world. For example, the waxy coating on cabbage leaves contains nonacosane ( $n\text{-C}_{29}\text{H}_{60}$ ), and the wood oil of the Jeffrey pine common to the Sierra Nevada contains heptane. Beeswax contains, among other things, hentriacontane ( $n\text{-C}_{31}\text{H}_{64}$ ).

By far the major sources of alkanes are the world's natural gas and petroleum deposits, laid down eons ago and derived from the decomposition of marine organic matter. **Natural gas** consists chiefly of methane, but ethane, propane, butane, and isobutane are also present. These simple

hydrocarbons are used in great quantities for heating homes, cooking food, and fueling industries. **Petroleum** is a complex mixture of hydrocarbons that must be refined into different fractions before it can be used.

**Refining** begins by fractional distillation of crude oil into three principal cuts: straight-run gasoline (bp 30–200°C), kerosene (bp 175–300°C), and gas oil (bp 275–400°C). Finally, distillation under reduced pressure gives lubricating oils and waxes and leaves an undistillable tarry residue of asphalt (Figure 3.4).



**Figure 3.4** The products of petroleum refining.

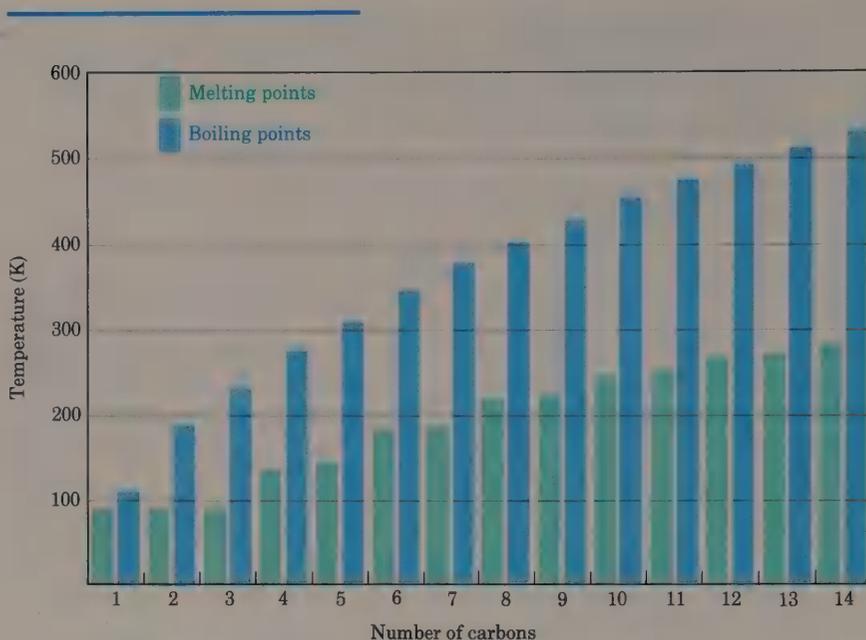
Simple distillation of crude oil is only the first step in gasoline production. It turns out that straight-run gasoline is a poor fuel because of the phenomenon of *engine knock*. In the ordinary automobile engine, a piston draws a mixture of fuel and air into a cylinder on its downward stroke and compresses the mixture on its upward stroke. Just before the end of the compression, a spark plug ignites the mixture and smooth combustion occurs, driving the piston downward.

Not all fuels burn equally well, though. When poor fuels are used, combustion can be initiated in an uncontrolled manner by a hot surface in the cylinder before the spark plug fires. This *preignition*, detected as an engine knock or ping, can destroy the engine by putting irregular forces on the crankshaft and by raising engine temperature.

The *octane number* of a fuel is the measure by which its antiknock properties are judged. It was recognized long ago that straight-chain hydrocarbons are far more prone to induce engine knock than are highly branched compounds. Heptane, a particularly bad fuel, is assigned a base value of 0 octane number; 2,2,4-trimethylpentane (commonly known as isooctane) has excellent antiknock characteristics and is given a rating of 100.

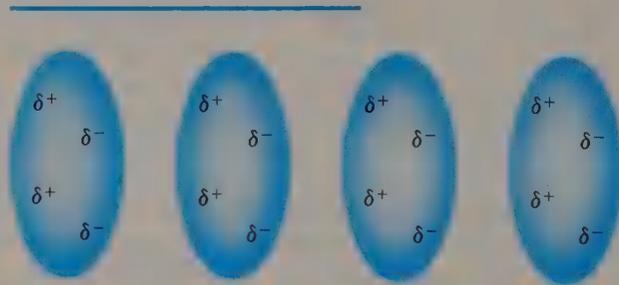






**Figure 3.5** A plot of melting and boiling points versus number of carbons for the  $C_1$ – $C_{15}$  alkanes. There is a regular increase with molecular size.

Nonpolar molecules like the alkanes are weakly attracted to each other by intermolecular **van der Waals forces**. These forces, which operate only over very small distances, result from induced polarization of the electron clouds in molecules. Although the electron distribution in a molecule is uniform on average over a period of time, the distribution at any given instant is not uniform. One side of a molecule may, by chance, have a slight excess of electrons relative to the opposite side. When that occurs, the molecule has a temporary dipole moment. This temporary dipole in one molecule causes a nearby molecule to adopt a temporarily opposite dipole, with the result that a tiny attraction is induced between the two molecules (Figure 3.6).



**Figure 3.6** Attractive van der Waals forces caused by temporary dipoles in molecules.

Temporary dipoles have a fleeting existence and are constantly changing, but the cumulative effect of an enormous number of these interactions produces attractive forces sufficient to cause the molecules to stay in the liquid state. Only when sufficient heat is applied to overcome these forces does the liquid boil.

As you might expect, van der Waals forces increase as molecule size increases. Although other factors are also involved, at least part of the increase in boiling point on going up the alkane series is due to increased van der Waals forces.

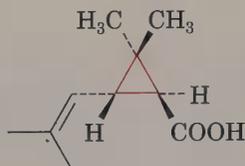
The effect of branching on boiling points can also be explained by invoking van der Waals forces. Branched alkanes are more nearly spherical than are straight-chain alkanes. As a result, they have smaller surface areas, smaller van der Waals forces, and consequently lower boiling points.

### 3.7 Cycloalkanes

Though we've discussed only open-chain alkanes up to this point, chemists have known for over 100 years that compounds with rings of carbon atoms also exist. Such compounds are called **cycloalkanes** or **alicyclic compounds (aliphatic cyclic)**. Since cycloalkanes consist of rings of  $-\text{CH}_2-$  units, they have the general formula  $(\text{CH}_2)_n$  or  $\text{C}_n\text{H}_{2n}$  and are represented by polygons in skeletal drawings:

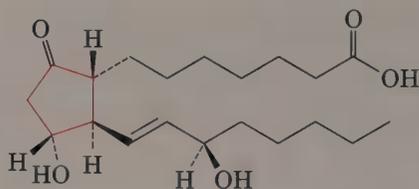


Alicyclic compounds with many different ring sizes abound in nature. For example, chrysanthemic acid contains a three-membered (cyclopropane) ring. Various esters of chrysanthemic acid occur naturally as the active insecticidal constituents of pyrethrum flowers.

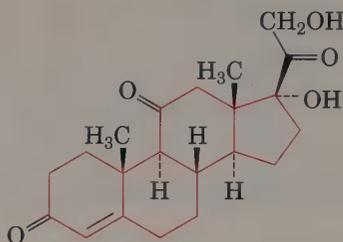


Chrysanthemic acid

*Prostaglandins*, such as  $\text{PGE}_1$ , contain a five-membered (cyclopentane) ring. Prostaglandins are potent hormones that control a wide variety of physiological functions in humans, including blood platelet aggregation, bronchial dilation, and inhibition of gastric secretions.

Prostaglandin E<sub>1</sub> (PGE<sub>1</sub>)

*Steroid hormones*, such as cortisone, contain four rings joined together—three six-membered (cyclohexane) and one five-membered (cyclopentane).



Cortisone

Some physical data for simple unsubstituted cycloalkanes are given in Table 3.6.

**TABLE 3.6** Physical Properties of Some Cycloalkanes

Name	Formula	Melting point (°C)	Boiling point (°C)	Density (g/mL)
Cyclopropane	C <sub>3</sub> H <sub>6</sub>	-127.6	-32.7	
Cyclobutane	C <sub>4</sub> H <sub>8</sub>	-50.0	12.0	0.720
Cyclopentane	C <sub>5</sub> H <sub>10</sub>	-93.9	49.3	0.7457
Cyclohexane	C <sub>6</sub> H <sub>12</sub>	6.6	80.7	0.7786
Cycloheptane	C <sub>7</sub> H <sub>14</sub>	-12.0	118.5	0.8098
Cyclooctane	C <sub>8</sub> H <sub>16</sub>	14.3	148.5	0.8349

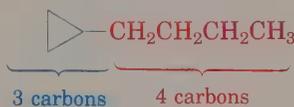
### 3.8 Naming Cycloalkanes

Cycloalkanes are named by rules similar to those used previously for open-chain alkanes. For most compounds, there are only two rules:

1. Use the cycloalkane name as the parent name. That is, compounds are normally named as alkyl-substituted cycloalkanes, rather than as cycloalkyl-substituted alkanes. The only exception to this rule occurs when the alkyl side chain contains a larger number of carbons than the ring. In such cases, the ring is considered a substituent on the parent open-chain alkane. For example:



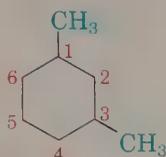
BUT



Methylcyclopentane

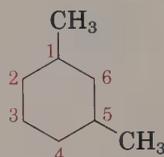
1-Cyclopropylbutane

2. Start at a point of attachment and number the substituents on the ring so as to arrive at the lowest sum:



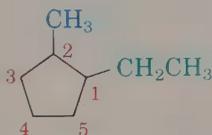
1,3-Dimethylcyclohexane

NOT



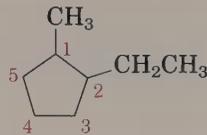
1,5-Dimethylcyclohexane

- a. When two or more different alkyl groups are present, they are numbered by alphabetical priority:



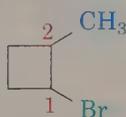
1-Ethyl-2-methylcyclopentane

NOT



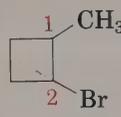
2-Ethyl-1-methylcyclopentane

- b. Halogen substituents, if present, are treated exactly like alkyl groups:



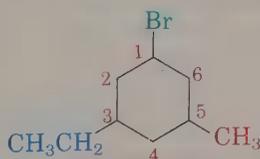
1-Bromo-2-methylcyclobutane

NOT

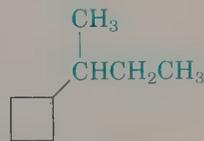
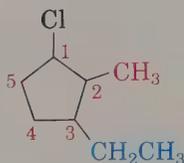


2-Bromo-1-methylcyclobutane

Some additional examples follow:



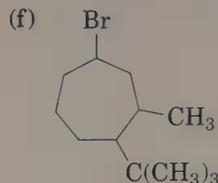
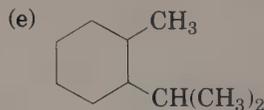
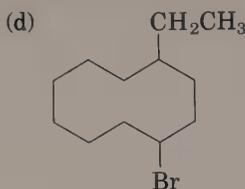
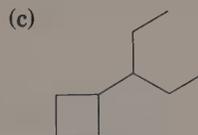
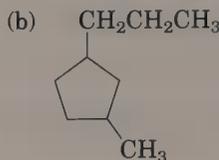
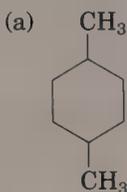
1-Bromo-3-ethyl-5-methylcyclohexane

(1-Methylpropyl)cyclobutane  
(or *sec*-butylcyclobutane)

1-Chloro-3-ethyl-2-methylcyclopentane

## PROBLEM .....

3.15 Give IUPAC names for the following cycloalkanes:



## PROBLEM .....

3.16 Draw structures corresponding to these IUPAC names:

(a) 1,1-Dimethylcyclooctane

(b) 3-Cyclobutylhexane

(c) 1,2-Dichlorocyclopentane

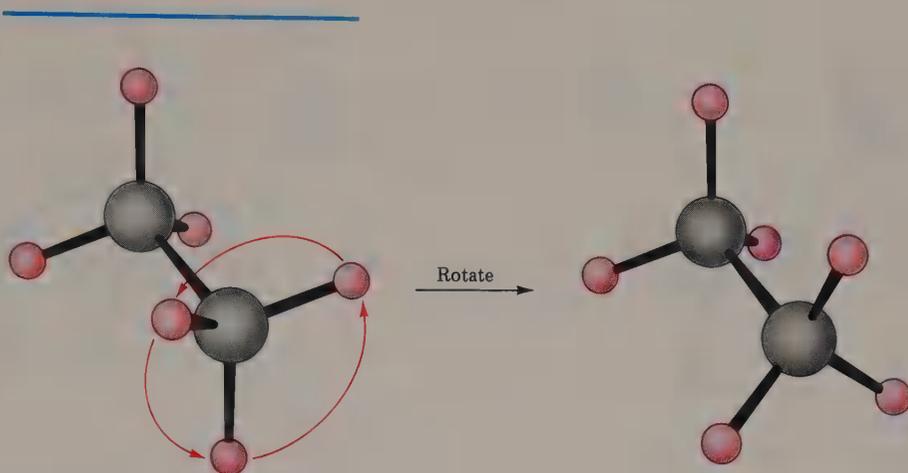
(d) 1,3-Dibromo-5-methylcyclohexane

### 3.9 Cis-Trans Isomerism in Cycloalkanes

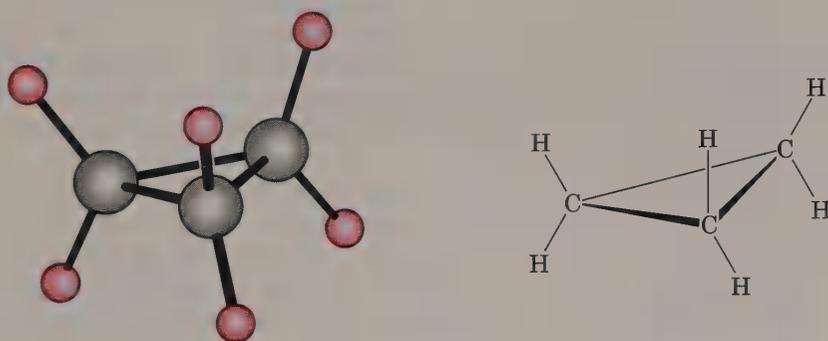
In most respects, the chemistry of cycloalkanes mimics that of open-chain, acyclic alkanes. Both classes of compounds are relatively nonpolar and are chemically inert to most reagents. There are, however, some important differences.

One difference is that cycloalkanes have less flexibility than their open-chain relatives. To see what this means, think for a minute about the nature of carbon-carbon single bonds. We know from Section 1.7 that sigma bonds are cylindrically symmetrical. In other words, the intersection of a plane cutting through a carbon-carbon single-bond orbital looks like a circle. Because of this cylindrical symmetry of sigma bonds, there is *free rotation* around carbon-carbon bonds in open-chain molecules. In ethane, for example, C-C bond overlap is exactly the same for all geometric arrangements of the hydrogens (Figure 3.7).

In contrast to the free rotation around single bonds in open-chain alkanes, there is much less freedom in cycloalkanes, which are geometrically constrained. Cyclopropane, for example, is geometrically constrained to be a flat, planar molecule with a rigid structure because three points define a plane. No bond rotation around a cyclopropane carbon-carbon bond can take place without breaking open the ring (Figure 3.8).



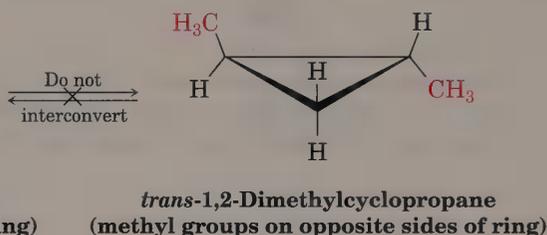
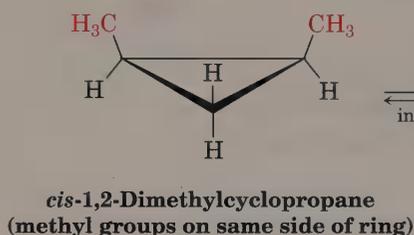
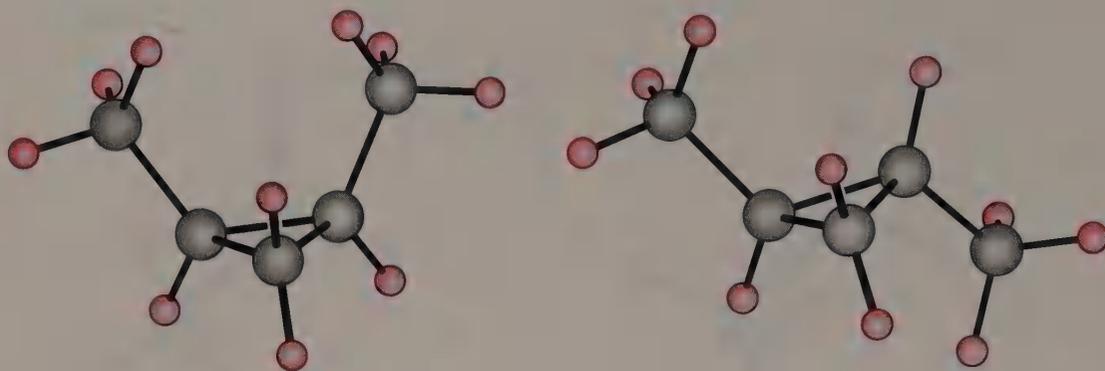
**Figure 3.7** Free rotation can occur around the carbon-carbon single bond in ethane because of sigma-bond cylindrical symmetry.



**Figure 3.8** The structure of cyclopropane. No rotation is possible around the carbon-carbon bonds without breaking open the ring.

Higher cycloalkanes have increasingly more freedom, and the very large rings ( $C_{25}$  and up) are so floppy that they are nearly indistinguishable from open-chain alkanes. The common ring sizes ( $C_3$ ,  $C_4$ ,  $C_5$ ,  $C_6$ ,  $C_7$ ), however, are restricted in their molecular motions.

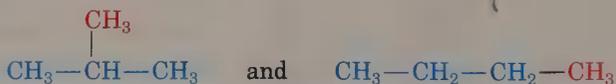
The most important consequence of their cyclic structure is that cycloalkanes have two sides, a "top" side and a "bottom" side, leading to the possibility of isomerism in substituted cycloalkanes. For example, there are two different 1,2-dimethylcyclopropane isomers, one with the two methyls on the same side of the ring and the other with them on opposite sides. Both isomers are stable compounds and can't be interconverted without breaking and re-forming chemical bonds. Make molecular models to prove this to yourself.



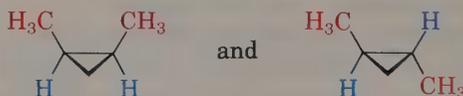
Do not  
interconvert

Unlike the *constitutional isomers* butane and isobutane (Section 3.3), which have their atoms connected in a different order, the two 1,2-dimethylcyclopropanes have the *same* order of connection but differ in the spatial orientation of their atoms. Such compounds that have their atoms connected in the same order but that differ in three-dimensional orientation are called **stereoisomers**.

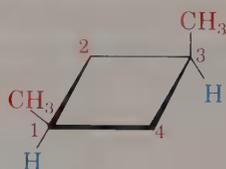
**Constitutional isomers**  
(different connections  
between atoms)



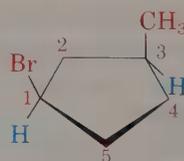
**Stereoisomers**  
(same connections but  
different three-dimensional  
geometry)



The 1,2-dimethylcyclopropanes are special kinds of stereoisomers called **cis–trans isomers**. The prefixes *cis*- (Latin; “on the same side”) and *trans*- (Latin; “across”) are used to distinguish between them. Cis–trans isomerism is a common occurrence in substituted cycloalkanes.



*cis*-1,3-Dimethylcyclobutane



*cis*-1-Bromo-3-methylcyclopentane

## PROBLEM .....

- 3.17 Draw the structures of these molecules:  
 (a) *trans*-1-Bromo-3-methylcyclohexane      (b) *cis*-1,2-Dimethylcyclopentane  
 (c) *trans*-1-*t*-Butyl-2-ethylcyclohexane
- .....

### 3.10 Summary and Key Words

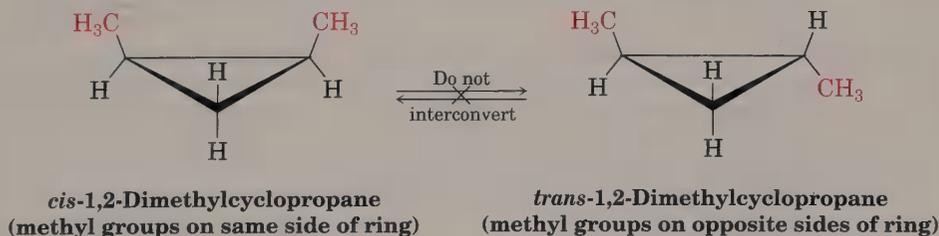
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A **functional group** is a part of a larger molecule. It is composed of an atom or a group of atoms that has characteristic chemical reactivity. Because functional groups behave approximately the same way in all molecules where they occur, the chemical reactions an organic molecule undergoes are largely determined by the functional groups it contains.

**Alkanes** are a class of **hydrocarbons** with the general formula  $C_nH_{2n+2}$ . They contain no functional groups, are chemically rather inert, and may be either straight-chain (**normal alkanes**) or branched. All alkanes can be uniquely named by a series of **IUPAC rules**.

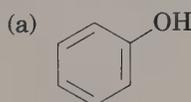
Compounds that have the same chemical formula but different structures are called **isomers**. More specifically, compounds such as butane and isobutane that differ in the order of their connections between atoms are called **constitutional isomers**.

**Cycloalkanes** contain rings of carbon atoms and have the general formula  $C_nH_{2n}$ . In comparison with open-chain alkanes, conformational mobility is greatly reduced in cycloalkanes. Although rotation is possible around carbon-carbon single bonds in open-chain molecules, full rotation around carbon-carbon bonds in cycloalkanes is not possible. Disubstituted cycloalkanes can therefore exist as **cis-trans isomers**. The *cis* isomer has both substituents on the same side of the ring; the *trans* isomer has substituents on opposite sides of the ring. *Cis-trans* isomers are just one kind of **stereoisomers**—isomers that have the same order of connection between atoms but that differ in their three-dimensional arrangements.

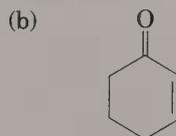


## ADDITIONAL PROBLEMS .....

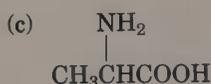
- 3.18 Locate and identify the functional groups in these molecules:



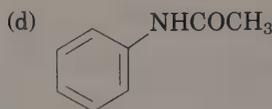
Phenol



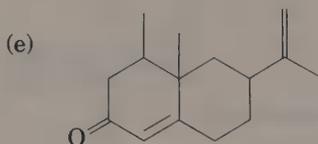
2-Cyclohexenone



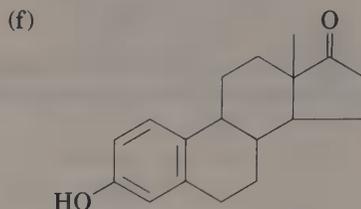
Alanine



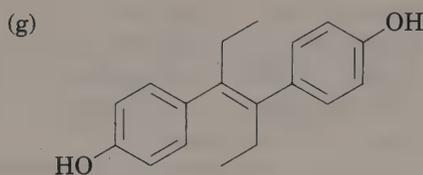
Acetanilide



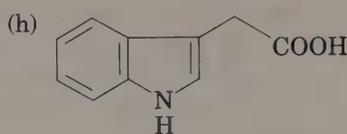
Nootkatone (from grapefruit)



Estrone



Diethylstilbestrol



3-Indoleacetic acid

**3.19** Draw structures that meet these descriptions (there are many possibilities):

- (a) Three isomers with the formula  $C_8H_{18}$   
 (b) Two isomers with the formula  $C_4H_8O_2$

**3.20** Draw structures of the nine isomers of  $C_7H_{16}$ .

**3.21** Propose structures that meet these descriptions:

- (a) A ketone with five carbons  
 (b) A four-carbon amide  
 (c) A five-carbon ester  
 (d) An aromatic aldehyde  
 (e) A keto ester  
 (f) An amino alcohol

**3.22** Propose suitable structures for the following:

- (a) A ketone,  $C_4H_8O$   
 (b) A nitrile,  $C_5H_9N$   
 (c) A dialdehyde,  $C_4H_6O_2$   
 (d) A bromoalkene,  $C_6H_{11}Br$   
 (e) An alkane,  $C_6H_{14}$   
 (f) A cycloalkane,  $C_6H_{12}$   
 (g) A diene (dialkene),  $C_5H_8$   
 (h) A keto alkene,  $C_5H_8O$

**3.23** How many compounds can you write that fit these descriptions?

- (a) Alcohols with formula  $C_4H_{10}O$   
 (b) Amines with formula  $C_5H_{13}N$   
 (c) Ketones with formula  $C_5H_{10}O$   
 (d) Aldehydes with formula  $C_5H_{10}O$   
 (e) Esters with formula  $C_4H_8O_2$   
 (f) Ethers with formula  $C_4H_{10}O$

**3.24** Draw compounds that contain the following:

- (a) A primary alcohol  
 (b) A tertiary nitrile  
 (c) A secondary bromide  
 (d) Both primary and secondary alcohols  
 (e) An isopropyl group  
 (f) A quaternary carbon

**3.25** Draw and name all monobromo derivatives of pentane,  $C_5H_{11}Br$ .

**3.26** Draw and name all monochloro derivatives of 2,5-dimethylhexane,  $C_8H_{17}Cl$ .

**3.27** What hybridization would you predict for the carbon atom in these functional groups?

- (a) Ketone  
 (b) Nitrile  
 (c) Carboxylic acid  
 (d) Ether

**3.28** Draw structural formulas for the following:

- (a) 2-Methylheptane  
 (b) 4-Ethyl-2,2-dimethylhexane  
 (c) 4-Ethyl-3,4-dimethyloctane  
 (d) 2,4,4-Trimethylheptane  
 (e) 3,3-Diethyl-2,5-dimethylnonane  
 (f) 4-Isopropyl-3-methylheptane

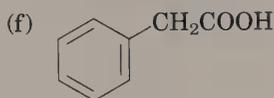
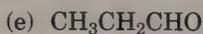
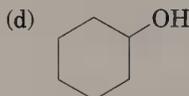
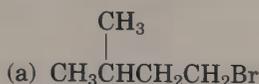
3.29 Draw a compound that:

- (a) Has only primary and tertiary carbons
- (b) Has no primary carbons
- (c) Has four secondary carbons

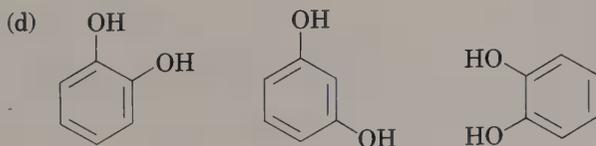
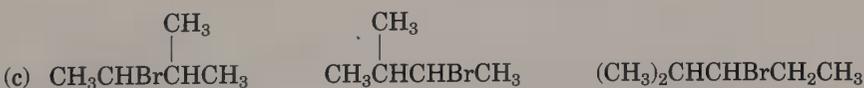
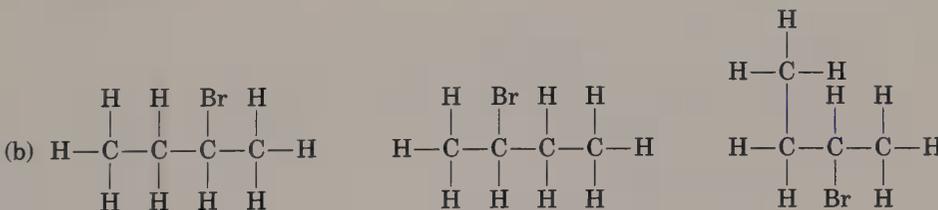
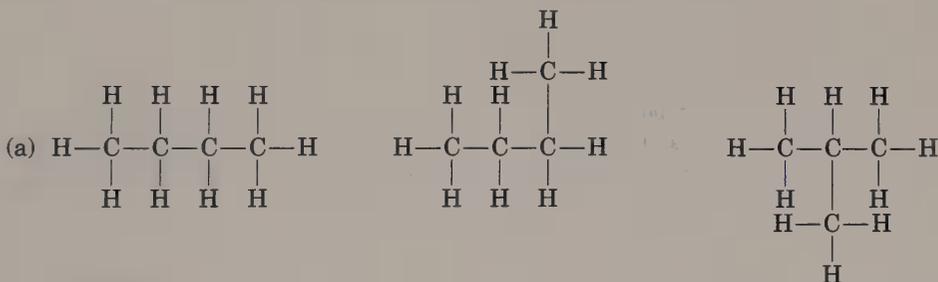
3.30 Draw a compound that:

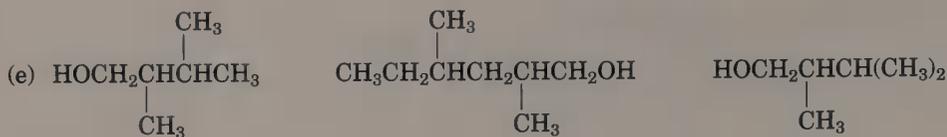
- (a) Has no primary hydrogens
- (b) Has only primary and tertiary hydrogens

3.31 For each of the following compounds, draw an isomer with the same functional groups:



3.32 Which of the following Kekulé structures represent the same compound, and which represent different compounds?





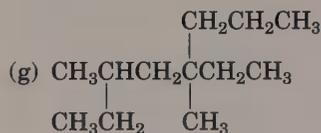
3.33 Draw structures for these compounds:

- (a) *trans*-1,3-Dibromocyclopentane                      (b) *cis*-1,4-Diethylcyclohexane  
 (c) *trans*-1-Isopropyl-3-methylcycloheptane      (d) Dicyclohexylmethane

3.34 Identify the kinds of carbons (1°, 2°, 3°, or 4°) in these molecules:



3.35 Supply proper IUPAC names for these compounds:



3.36 Name the five isomers of  $\text{C}_6\text{H}_{14}$ .

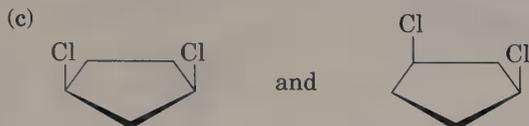
3.37 The following names are incorrect. Give the proper IUPAC names.

- (a) 2,2-Dimethyl-6-ethylheptane                      (b) 4-Ethyl-5,5-dimethylpentane  
 (c) 3-Ethyl-4,4-dimethylhexane                      (d) 5,5,6-Trimethyloctane  
 (e) 2-Isopropyl-4-methylheptane                      (f) *cis*-1,5-Dimethylcyclohexane

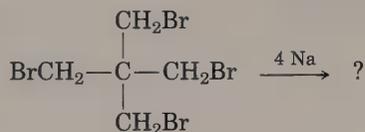
3.38 Propose structures and give the correct IUPAC names for the following:

- (a) A dimethylcyclooctane                      (b) A diethyldimethylhexane  
 (c) A cyclic alkane with three methyl groups  
 (d) A (3-methylbutyl)-substituted alkane

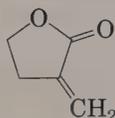
- 3.39 Draw 1,3,5-trimethylcyclohexane using a hexagon to represent the ring. How many cis-trans stereoisomers are possible?
- 3.40 Tell whether the following pairs of compounds are identical, constitutional isomers, or stereoisomers.
- (a) *cis*-1,3-Dibromocyclohexane and *trans*-1,4-dibromocyclohexane
- (b) 2,3-Dimethylhexane and 2,5,5-trimethylpentane



- 3.41 Draw two constitutional isomers of *cis*-1,2-dibromocyclopentane.
- 3.42 Draw a stereoisomer of *trans*-1,3-dimethylcyclobutane.
- 3.43 Malic acid, a compound of formula  $C_4H_6O_5$ , has been isolated from apples. Since this compound reacts with 2 equivalents (equiv) of base, it can be formulated as a dicarboxylic acid.
- (a) Draw at least five possible structures.
- (b) This compound can be shown to be a secondary alcohol. What is the structure of malic acid?
- 3.44 Cyclopropane was first prepared by reaction of 1,3-dibromopropane with sodium metal. Formulate the cyclopropane-forming reaction and then predict the product that you might obtain from the following reaction. What geometry would you expect for the product?

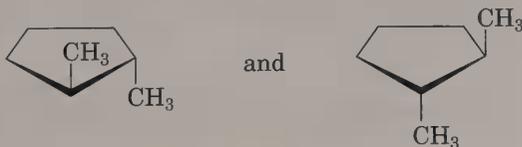


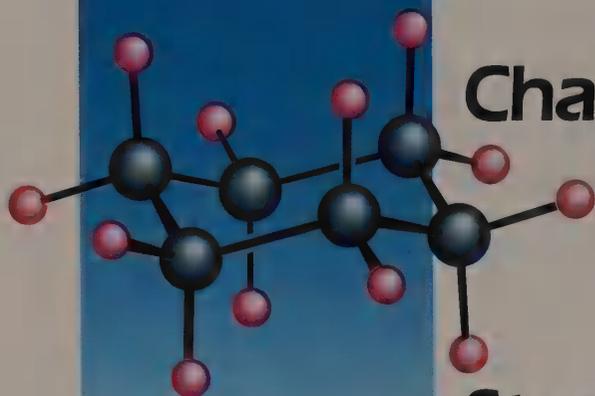
- 3.45 The compound  $\alpha$ -methylenebutyrolactone is a skin irritant that has been isolated from the dogtooth violet. What functional groups does it contain?



$\alpha$ -Methylenebutyrolactone

- 3.46 Formaldehyde,  $H_2C=O$ , is a simple compound known to all biologists because of its usefulness as a tissue preservative. When pure, formaldehyde *trimerizes* to give trioxane,  $C_3H_6O_3$ . Trioxane, surprisingly enough, has no carbonyl groups. Only one monobromo derivative of trioxane is possible. Propose a structure that fits these data.
- 3.47 Here's a tough one. There are two different substances named *trans*-1,2-dimethylcyclopentane. Make molecular models and see if you can find the relationship between them. We'll explore this kind of isomerism in Chapter 9.





## Chapter 4

# Stereochemistry of Alkanes and Cycloalkanes

Up to this point, we've viewed most molecules in a two-dimensional way and have given little thought to any chemical consequences that might arise from the spatial arrangement of atoms in molecules. Now it's time to add a third dimension to our study. **Stereochemistry** is the branch of chemistry concerned with the three-dimensional aspects of molecules.

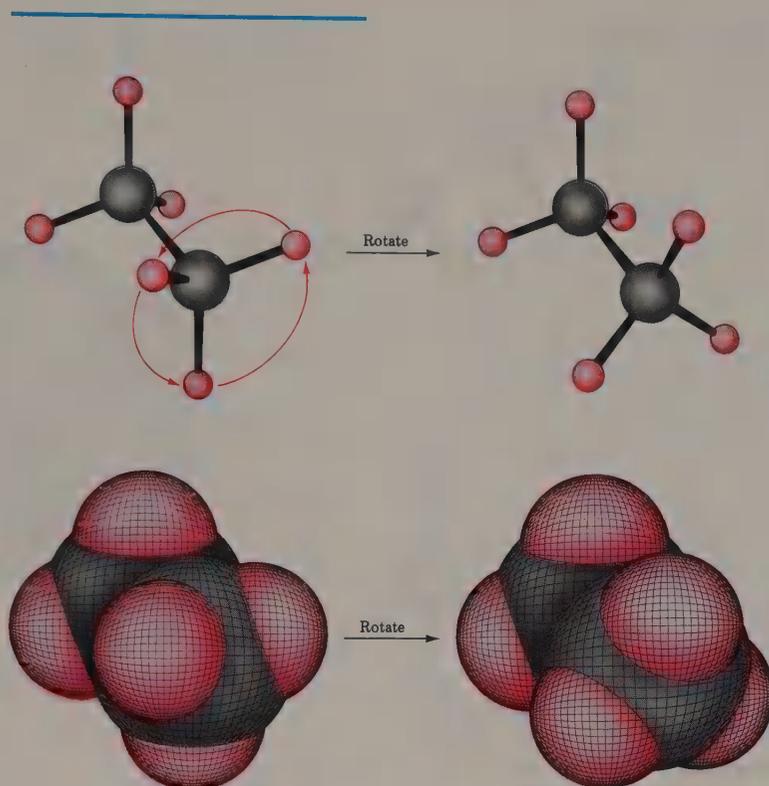
### 4.1 Conformations of Ethane

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We know that an  $sp^3$ -hybridized carbon atom has tetrahedral geometry and that the carbon-carbon bonds in alkanes result from sigma overlap of two carbon  $sp^3$  orbitals. Let's now look into the three-dimensional consequences of such bonding. What are the spatial relationships between the hydrogens on one carbon and the hydrogens on a neighboring carbon?

We saw in Section 3.9 that free rotation can occur around carbon-carbon single bonds in open-chain molecules such as ethane, as a consequence of sigma bond cylindrical symmetry. Orbital overlap in the C-C single bond is exactly the same for all geometric arrangements of the atoms (Figure 4.1). The different arrangements of atoms caused by rotation about a single bond are called **conformations**, and a specific conformation is called a **conformer** (**conformational isomer**). Unlike constitutional isomers, though, different conformers can't usually be isolated because they interconvert too rapidly.

Chemists represent conformational isomers in two ways, as shown in Figure 4.2. **Sawhorse representations** view the carbon-carbon bond from an oblique angle and indicate spatial orientation by showing all of the C-H

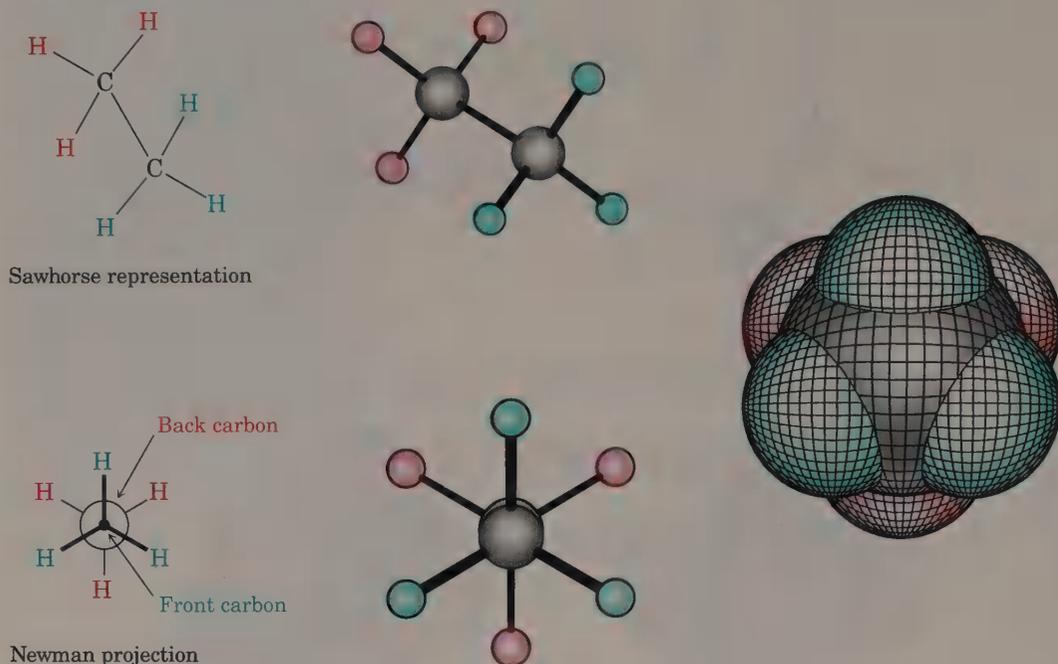


**Figure 4.1** Some conformations of ethane. Rapid rotation around the carbon–carbon single bond interconverts the different forms. The drawings are computer generated.

bonds. **Newman<sup>1</sup> projections** view the carbon–carbon bond directly end-on and represent the two carbon atoms by a circle. Bonds attached to the front carbon are represented by lines going to the center of the circle, and bonds attached to the rear carbon are represented by lines going to the edge of the circle. The advantages of Newman projections are that they're easy to draw and the relationships among substituents on the different carbon atoms are easily seen.

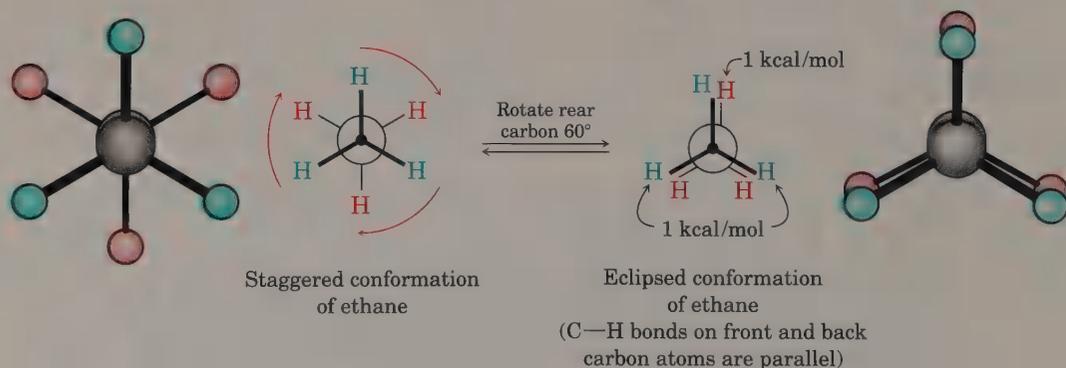
In spite of what we've just said about sigma bond symmetry, we don't actually observe *perfectly* free rotation in ethane. Experiments show that there is a slight (2.9 kcal/mol; 12 kJ/mol) barrier to rotation and that some conformations are more stable (have less energy) than others. The lowest-energy, most stable conformation is the one in which all six carbon–hydrogen bonds are as far away from each other as possible (**staggered** when viewed end-on in a Newman projection). The highest-energy, least stable confor-

<sup>1</sup>Melvin S. Newman (1908– ); b. New York; Ph.D. (1932), Yale University; professor, Ohio State University.



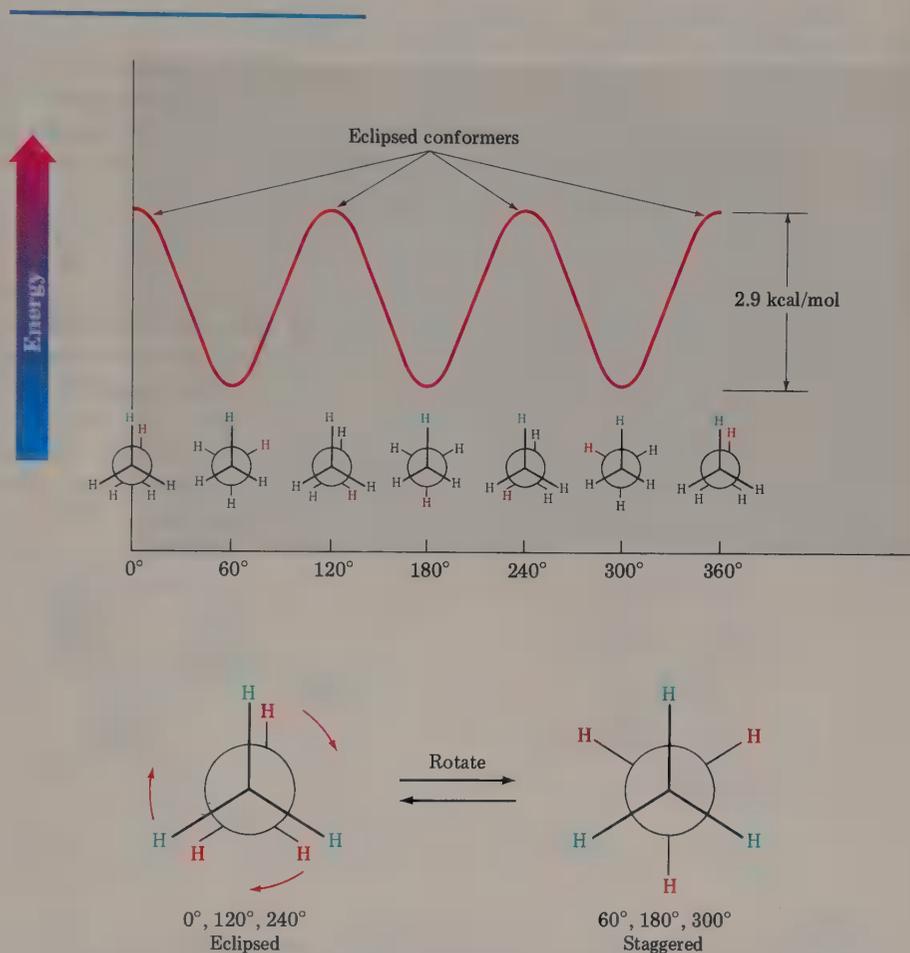
**Figure 4.2** A sawhorse representation and a Newman projection of ethane.

mation is the one in which the six carbon–hydrogen bonds are as close as possible (**eclipsed** in a Newman projection). In between these two limiting conformations are, of course, an infinite number of other possibilities. Since the barrier to rotation is 2.9 kcal/mol, and since the barrier is caused by three equal hydrogen–hydrogen eclipsing interactions, we can assign a value of approximately 1 kcal/mol (4 kJ/mol) to each single interaction.



The 2.9 kcal/mol of extra energy present in the eclipsed conformation of ethane is called **torsional strain**. The barrier to rotation that results from torsional strain can be represented on a graph of potential energy

versus degree of rotation, where the angle between C–H bonds on front and back carbons (the *dihedral angle*) goes full circle from  $0^\circ$  to  $360^\circ$  when seen end-on. Energy minima occur at staggered conformations, and energy maxima occur at eclipsed conformations, as shown in Figure 4.3.



**Figure 4.3** A graph of potential energy versus bond rotation in ethane. The staggered conformers are 2.9 kcal/mol lower in energy than the eclipsed conformers.

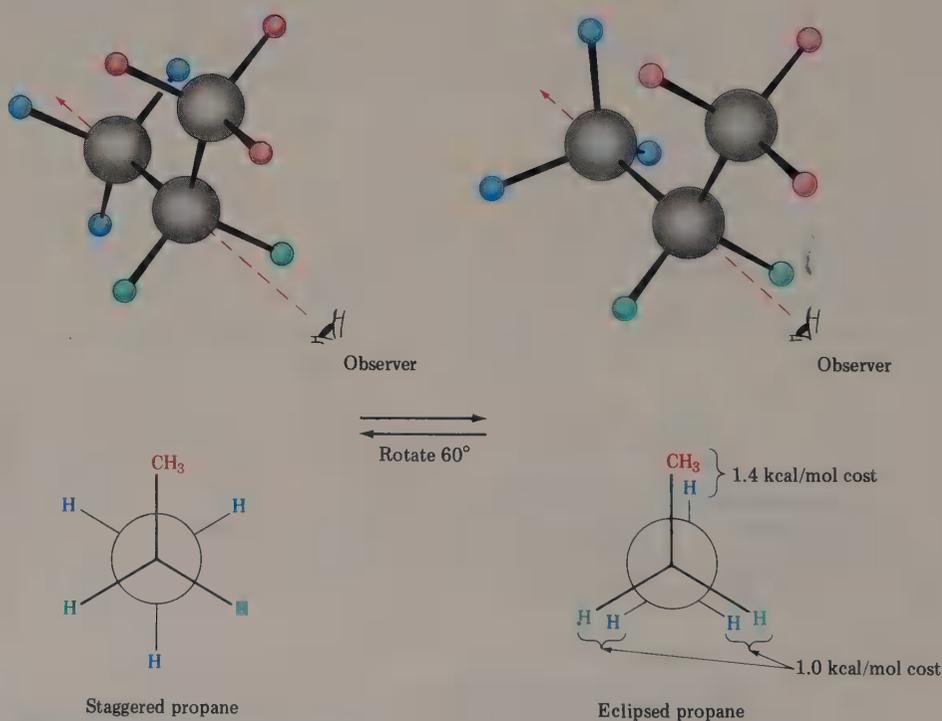
To what is torsional strain due? The reasons for its existence have been the subject of some controversy, but most chemists now believe that torsional strain is due to the slight repulsion between electron clouds in the carbon–hydrogen bonds as they pass by each other at close quarters in the eclipsed conformer. Calculations indicate that the internuclear hydrogen–hydrogen distance in the staggered conformer is  $2.55 \text{ \AA}$  but that this distance decreases to about  $2.29 \text{ \AA}$  in the eclipsed conformer.

## PROBLEM .....

- 4.1 Build a molecular model of ethane, and look at the interconversion of staggered and eclipsed forms.
- .....

## 4.2 Conformations of Propane

Propane is the next higher member in the alkane series, and we again find a torsional barrier that results in slightly hindered rotation around the carbon-carbon bonds. The barrier is slightly higher in propane than in ethane—3.4 kcal/mol (14 kJ/mol) versus 2.9 kcal/mol. In the eclipsed conformer of propane, there are two ethane-type hydrogen-hydrogen interactions and one additional interaction between a carbon-hydrogen bond and a carbon-carbon bond. Since each eclipsing hydrogen-hydrogen interaction has an energy “cost” of 1.0 kcal/mol, we can assign a value of  $3.4 - (2 \times 1.0 \text{ kcal/mol}) = 1.4 \text{ kcal/mol}$  (6 kJ/mol) to the interaction between the carbon-methyl bond and the carbon-hydrogen bond (Figure 4.4).



**Figure 4.4** Newman projections of propane showing staggered and eclipsed conformations. The staggered conformer is lower in energy by 3.4 kcal/mol.

## PROBLEM.....

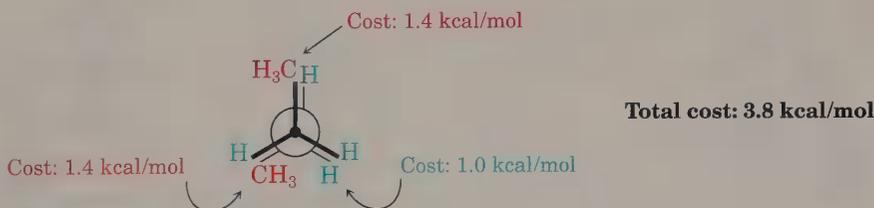
- 4.2 Make a graph of potential energy versus angle of bond rotation for propane. Assign quantitative values to the energy maxima.
- .....

## 4.3 Conformations of Butane

---

The conformational situation is more complex for higher alkanes. In butane, for instance, a plot of potential energy versus rotation about the C2–C3 bond is shown in Figure 4.5.

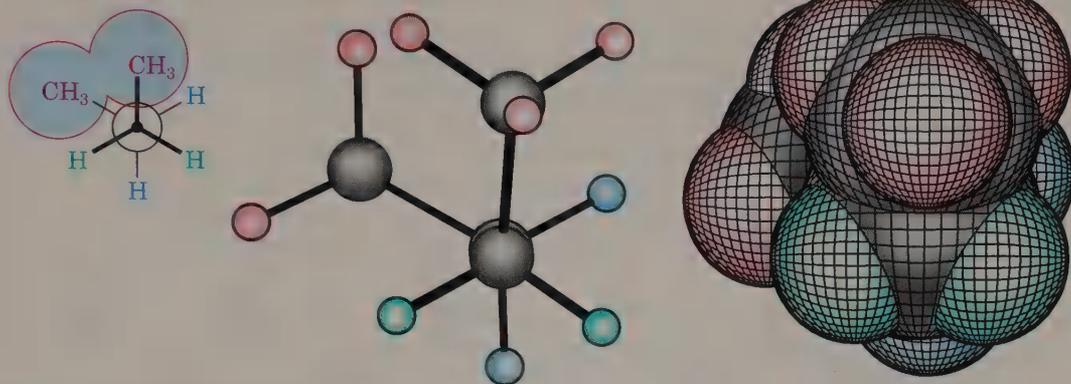
Not all staggered conformations of butane have the same energy, and not all eclipsed conformations are the same. The lowest-energy arrangement, called the **anti conformation**, is the one in which the two large groups (methyls) are as far apart as possible—that is,  $180^\circ$ . As rotation around the C2–C3 bond occurs, an eclipsed conformation is reached in which there are two methyl–hydrogen interactions and one hydrogen–hydrogen interaction. If we assign the energy values for eclipsing interactions that were previously derived from ethane and propane, we predict that this eclipsed conformation should be more strained than the anti conformation by  $2 \times 1.4$  kcal/mol (two methyl–hydrogen interactions) plus 1.0 kcal/mol (one hydrogen–hydrogen interaction), or a total of 3.8 kcal/mol (16 kJ/mol). This is exactly what is observed.



When bond rotation is continued, an energy minimum is reached at the staggered conformation where the methyl groups are  $60^\circ$  apart. Called the **gauche conformation**, it lies 0.9 kcal/mol higher in energy than the anti conformation *even though it has no eclipsing interactions*. This energy difference is due to the fact that the large methyl groups are near each other in the gauche conformation, resulting in *steric strain*. **Steric strain** is the repulsive interaction that occurs when two groups are forced to be closer to each other than their atomic radii allow. Basically, it's the result of trying to force two atoms to occupy the same space (Figure 4.6).

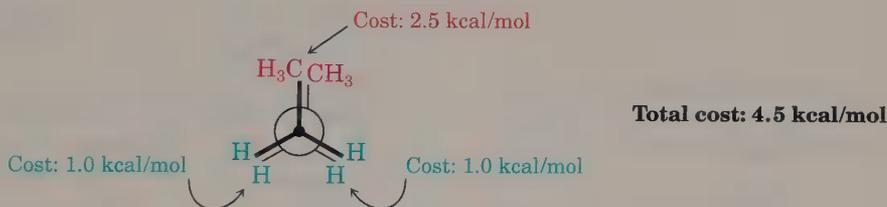
As the dihedral angle between the methyl groups approaches  $0^\circ$ , the energy maximum is reached. Since the methyl groups are forced even closer together than in the gauche conformation, a large amount of both torsional and steric strain is present. A total strain energy of 4.5 kcal/mol (19 kJ/mol) has been estimated for this conformation, allowing us to calculate a





**Figure 4.6** The spatial interaction between two methyl groups in gauche butane. Steric strain results when the two methyl groups are too close together.

value of 2.5 kcal/mol (11 kJ/mol) for the methyl–methyl eclipsing interaction [total strain (4.5 kcal/mol), less strain of two hydrogen–hydrogen eclipsing interactions ( $2 \times 1.0$  kcal/mol), results in 2.5 kcal/mol].



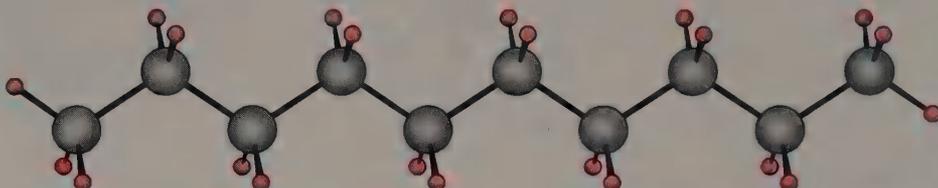
After  $0^\circ$ , the rotation becomes a mirror image of what we've already seen. Another gauche conformation is reached, another eclipsed conformation, and finally a return to the anti conformation occurs.

The concept of assigning definite energy values to specific interactions within a molecule is a very useful one that we'll return to later in this chapter. A summary of what we've found thus far is given in Table 4.1.

**TABLE 4.1** Energy Costs for Interactions in Alkane Conformers

Interaction	Cause	Energy cost	
		(kcal/mol)	(kJ/mol)
H—H eclipsed	Torsional strain	1.0	4
H—CH <sub>3</sub> eclipsed	Mostly torsional strain	1.4	6
CH <sub>3</sub> —CH <sub>3</sub> eclipsed	Torsional plus steric strain	2.5	11
CH <sub>3</sub> —CH <sub>3</sub> gauche	Steric strain	0.9	4

The same principles just developed for butane apply to pentane, hexane, and all higher alkanes. The most favored conformation for any alkane is the one in which all carbon-carbon bonds have staggered arrangements and in which large substituents are arranged anti to each other. A generalized alkane structure is shown in Figure 4.7.



**Figure 4.7** The most stable alkane conformation is the one in which all substituents are staggered and in which the carbon-carbon bonds are arranged anti, as in this computer-generated structure of decane.

One final point: It's important to realize that when we speak of a particular conformer as being "more stable" than another, we don't mean the molecule adopts and maintains only the more stable conformation. At room temperature, enough thermal energy is present to ensure that rotation around sigma bonds occurs rapidly and that all conformers are in a fluid equilibrium. At any given instant, however, a larger percentage of molecules will be found in a more stable conformation than in a less stable one.

**PROBLEM** .....

- 4.3 Sight along the C2-C3 bond of 2,3-dimethylbutane, and draw a Newman projection of the most stable conformation.

**PROBLEM** .....

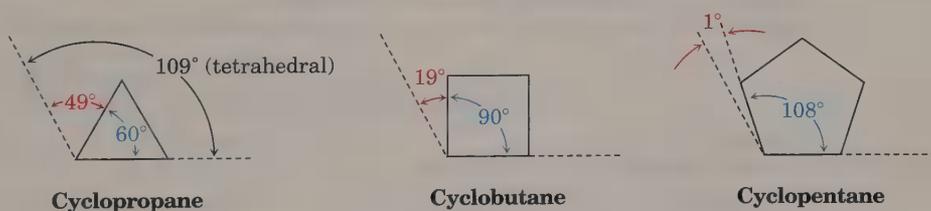
- 4.4 Consider 2-methylpropane (isobutane). Sighting along the C2-C1 bond:
- Draw a Newman projection of the most stable conformation.
  - Draw a Newman projection of the least stable conformation.
  - Make a qualitative graph of energy versus angle of rotation about the C2-C1 bond.
  - Since a hydrogen-hydrogen eclipsing interaction "costs" 1.0 kcal/mol and a hydrogen-methyl eclipsing interaction costs 1.4 kcal/mol, assign quantitative values to your graph.
- .....

## 4.4 Conformation and Stability of Cycloalkanes: The Baeyer Strain Theory

Chemists in the late 1800s had accepted the idea that cyclic molecules existed, but the limitations on feasible ring sizes were unclear. Numerous compounds containing five-membered and six-membered rings were known, but smaller and larger ring sizes had not been prepared. For example, no cyclopropanes or cyclobutanes were known, despite numerous efforts to prepare them.

A theoretical interpretation of this observation was proposed in 1885 by Adolf von Baeyer.<sup>2</sup> Baeyer suggested that, if carbon prefers to have tetrahedral geometry with bond angles of  $109^\circ$ , ring sizes other than five and six may be too *strained* to exist. Baeyer based his hypothesis on the simple geometric notion that a three-membered ring (cyclopropane) should be an equilateral triangle with bond angles of  $60^\circ$ , a four-membered ring (cyclobutane) should be a square with bond angles of  $90^\circ$ , a five-membered ring (cyclopentane) should be a regular pentagon with bond angles of  $108^\circ$ , and so on.

According to Baeyer, cyclopropane, with a bond angle compression of  $109^\circ - 60^\circ = 49^\circ$ , should have a large amount of *angle strain* and must therefore be highly reactive. Cyclobutane ( $109^\circ - 90^\circ = 19^\circ$  angle strain) must be similarly reactive, but cyclopentane ( $109^\circ - 108^\circ = 1^\circ$  angle strain) must be nearly strain-free. Cyclohexane ( $109^\circ - 120^\circ = -11^\circ$  angle strain) must be somewhat strained, but cycloheptane ( $109^\circ - 128^\circ = -19^\circ$  angle strain) and higher cycloalkanes must have bond angles that are forced to be too large. Carrying this line of reasoning further, Baeyer suggested that very large rings should be impossibly strained and incapable of existence.



Although there is some truth to Baeyer's assertions about angle strain in small rings, he was incorrect in his belief that small and large rings are too strained to exist. Rings of all sizes from 3 through 30 and beyond can now be prepared. Nevertheless, the concept of **angle strain**—the resistance of a bond angle to compression or expansion from the ideal tetrahedral angle—is a very useful one. Let's see what the facts are.

<sup>2</sup>Adolf von Baeyer (1835–1917); b. Berlin; Ph.D. Berlin (1858); professor, Berlin, Strasbourg (1872–1875), Munich (1875–1917); Nobel Prize (1905).

## 4.5 Heats of Combustion of Cycloalkanes

How can the amount of strain in a compound be measured? We must first measure the total amount of energy in the compound and then subtract the amount of energy in a hypothetical strain-free reference compound. The difference between the two values should represent the amount of extra energy in the molecule due to strain.

The simplest way to determine strain energies is to measure heats of combustion of the cycloalkanes. The **heat of combustion** of a compound is the amount of heat (energy) released when the compound burns completely with oxygen:

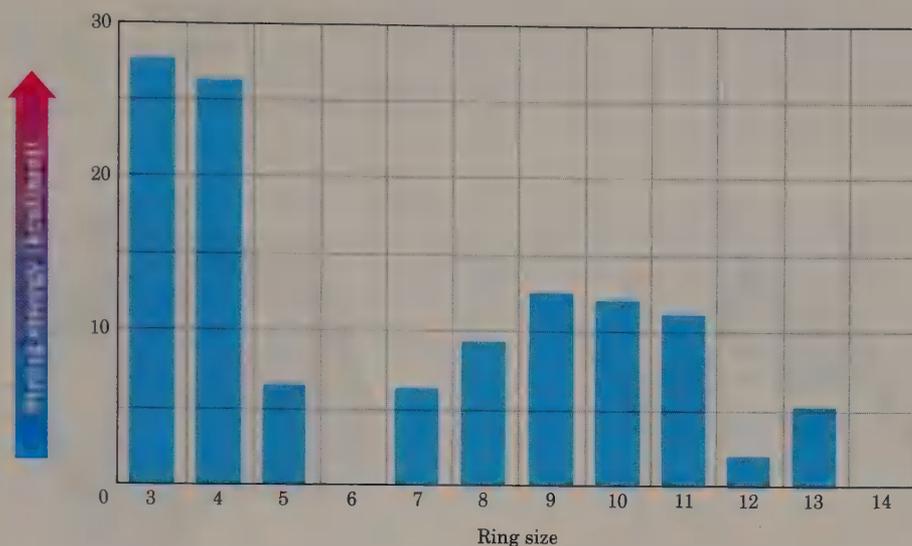


The more energy (strain) the sample contains, the more energy (heat) that's released on combustion. If we compare the heats of combustion of two isomeric substances, more energy is released during combustion of the more strained substance because that compound has a higher energy level to begin with.

Since the heat of combustion of a hydrocarbon depends on its molecular weight, it's more useful for comparison purposes to look at heats of combustion per  $\text{CH}_2$  unit. In this way, the size of the hydrocarbon is not a factor, and we can compare cycloalkane rings of different sizes to a standard, strain-free, acyclic alkane. Table 4.2 and Figure 4.8 show the results of this comparison. Total strain energies are calculated by taking the difference between sample heat of combustion per  $\text{CH}_2$  and reference heat of combustion per  $\text{CH}_2$ , and multiplying by the number of carbons,  $n$ , in the sample ring.

**TABLE 4.2** Heats of Combustion of Cycloalkanes

Cycloalkane $(\text{CH}_2)_n$	Ring size, $n$	Heat of combustion (kcal/mol)	Heat of combustion per $\text{CH}_2$ (kcal/mol)	Total strain energy (kcal/mol)
Cyclopropane	3	499.8	166.6	27.6
Cyclobutane	4	655.9	164.0	26.4
Cyclopentane	5	793.5	158.7	6.5
Cyclohexane	6	944.5	157.4	0
Cycloheptane	7	1108	158.3	6.3
Cyclooctane	8	1269	158.6	9.6
Cyclononane	9	1429	158.8	12.6
Cyclodecane	10	1586	158.6	12.0
Cycloundecane	11	1742	158.4	11.0
Cyclododecane	12	1891	157.6	2.4
Cyclotridecane	13	2051	157.8	5.2
Cyclotetradecane	14	2204	157.4	0
Alkane (reference)			157.4	0



**Figure 4.8** Cycloalkane strain energy as a function of ring size. Note that cyclohexane rings are strain-free.

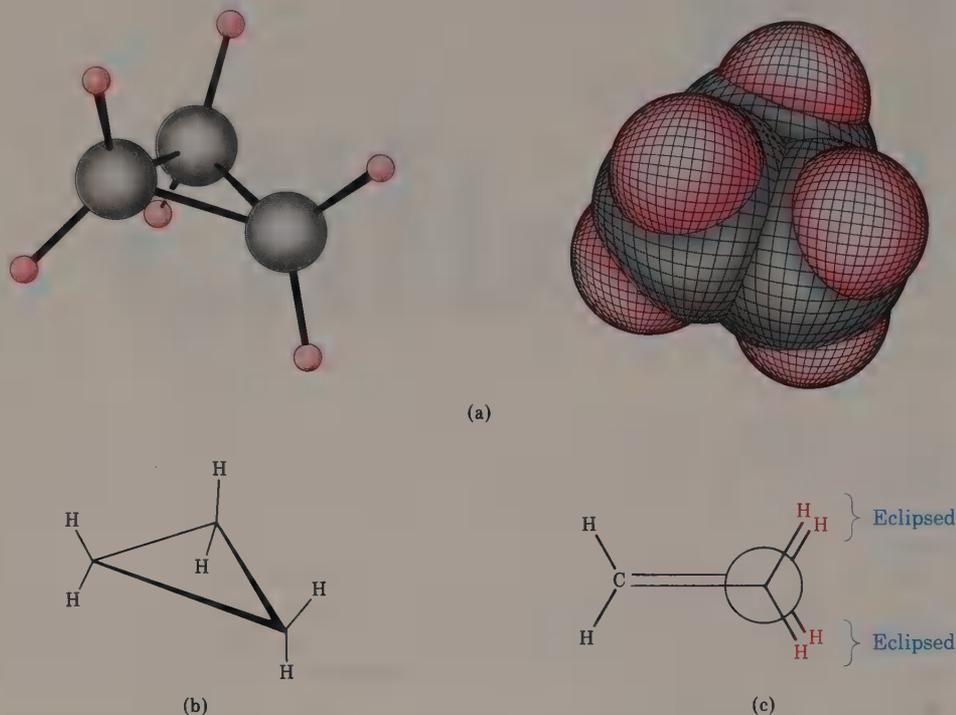
The data in Table 4.2 and the graph in Figure 4.8 show clearly that Baeyer's theory is not fully correct. Cyclopropane and cyclobutane are indeed quite strained, just as predicted, but cyclopentane is more strained than predicted, and cyclohexane is strain-free. For rings of larger size, there is no regular increase in strain, and rings having more than 14 carbons are again strain-free. Why is Baeyer's theory wrong?

## 4.6 The Nature of Ring Strain

Baeyer was wrong for a very simple reason: He assumed that all rings are flat. In fact, though, most cycloalkanes *aren't* flat; they adopt puckered three-dimensional conformations that allow bond angles to be nearly tetrahedral. Nevertheless, the concept of angle strain goes far toward explaining the reactivity of three- and four-membered rings.

Several factors in addition to angle strain are important in determining the shape and total strain energy of rings. One factor is the resistance to bond rotation, or torsional strain, encountered earlier in the discussion of alkane conformations (Section 4.1). We said at that time that open-chain alkanes are most stable in a staggered conformation and least stable in an eclipsed conformation. A similar conclusion holds for cycloalkanes: Torsional strain is present in cycloalkanes unless all of the bonds have a staggered arrangement. For example, cyclopropane must have considerable torsional strain (in addition to angle strain), because C-H bonds on neighboring

carbon atoms are eclipsed (Figure 4.9). Larger cycloalkanes minimize this strain by adopting puckered, nonplanar conformations.



**Figure 4.9** The conformation of cyclopropane. Part (c) is a Newman projection along a C–C bond, showing the eclipsing of neighboring C–H bonds.

Steric strain (Section 4.3) is a third factor that contributes to the overall strain energy of cycloalkanes. As we saw in gauche butane (Section 4.3), two nonbonded groups repel each other if they approach too closely and attempt to occupy the same space. Such nonbonded steric interactions are particularly important in determining the minimum-energy conformations of medium-ring ( $C_7$ – $C_{11}$ ) cycloalkanes.

In summary, the Baeyer theory is insufficient to explain the observed strain energies and geometries of cycloalkanes. Cycloalkanes adopt their minimum-energy conformations for a combination of three reasons:

1. **Angle strain**, the strain due to expansion or compression of bond angles
2. **Torsional strain**, the strain due to eclipsing of neighboring bonds
3. **Steric strain**, the strain due to repulsive interactions when atoms approach each other too closely

## PROBLEM .....

- 4.5 We saw in Section 4.1 that each hydrogen–hydrogen eclipsing interaction in the eclipsed conformation of ethane costs about 1.0 kcal/mol. How many such eclipsing interactions are present in cyclopropane? What fraction of the overall 27.6 kcal/mol (115 kJ/mol) strain energy of cyclopropane is due to torsional strain?

## PROBLEM .....

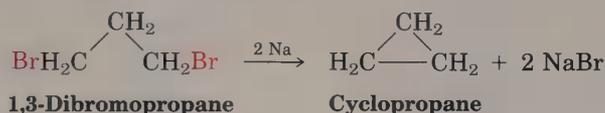
- 4.6 *cis*-1,2-Dimethylcyclopropane has a higher heat of combustion than *trans*-1,2-dimethylcyclopropane. How can you account for this difference? Which of the two compounds is more stable?

## PROBLEM .....

- 4.7 If both propane and cyclopropane were equally available and equally priced, which would be the more efficient fuel? Explain.

## 4.7 Cyclopropane: An Orbital View

Cyclopropane is a colorless gas (bp =  $-33^{\circ}\text{C}$ ) that was first prepared by reaction of sodium with 1,3-dibromopropane:

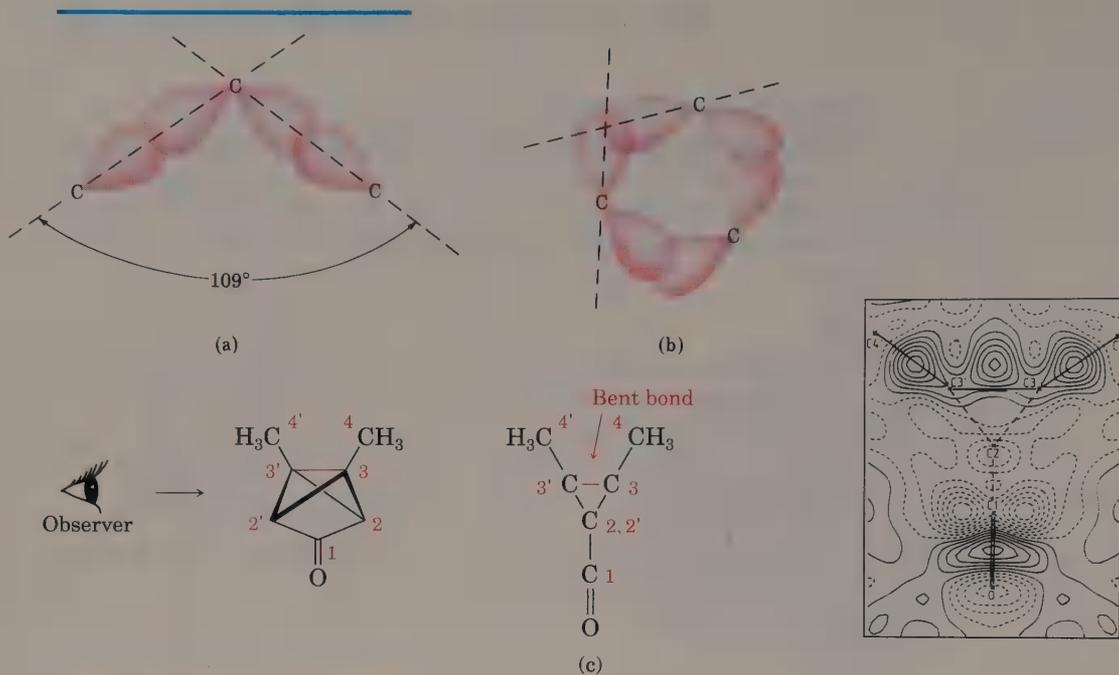


Since three points (the carbon atoms) define a plane, cyclopropane must be flat and, assuming it's symmetrical, must have C–C–C bond angles of  $60^{\circ}$ . How can molecular orbital theory account for this great distortion of the bonds from the normal  $109^{\circ}$  tetrahedral angle?

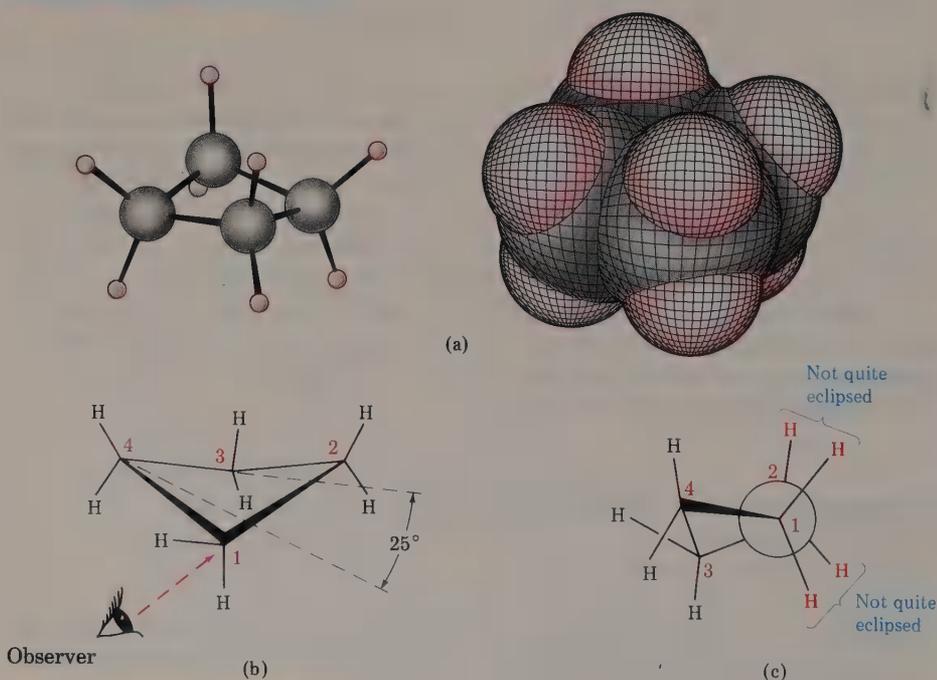
The answer is that cyclopropane is best thought of as having **bent bonds**. In an unstrained alkane, maximum bonding is achieved when two atoms are located so that their overlapping orbitals point directly toward each other. In cyclopropane, though, the orbitals cannot point directly toward each other; rather, they must overlap at a slight angle (Figure 4.10). The result of this poor overlap is that cyclopropane bonds are weaker and more reactive than normal alkane bonds.

## 4.8 Conformations of Cyclobutane and Cyclopentane

Cyclobutane has nearly the same total amount of strain [26.4 kcal/mol (110 kJ/mol)] as cyclopropane [27.6 kcal/mol (115 kJ/mol)]. Because of its larger number of ring hydrogens, cyclobutane has more torsional strain than cyclopropane has even though it has less angle strain. Spectroscopic measurements indicate that cyclobutane is not quite flat but is slightly bent so that



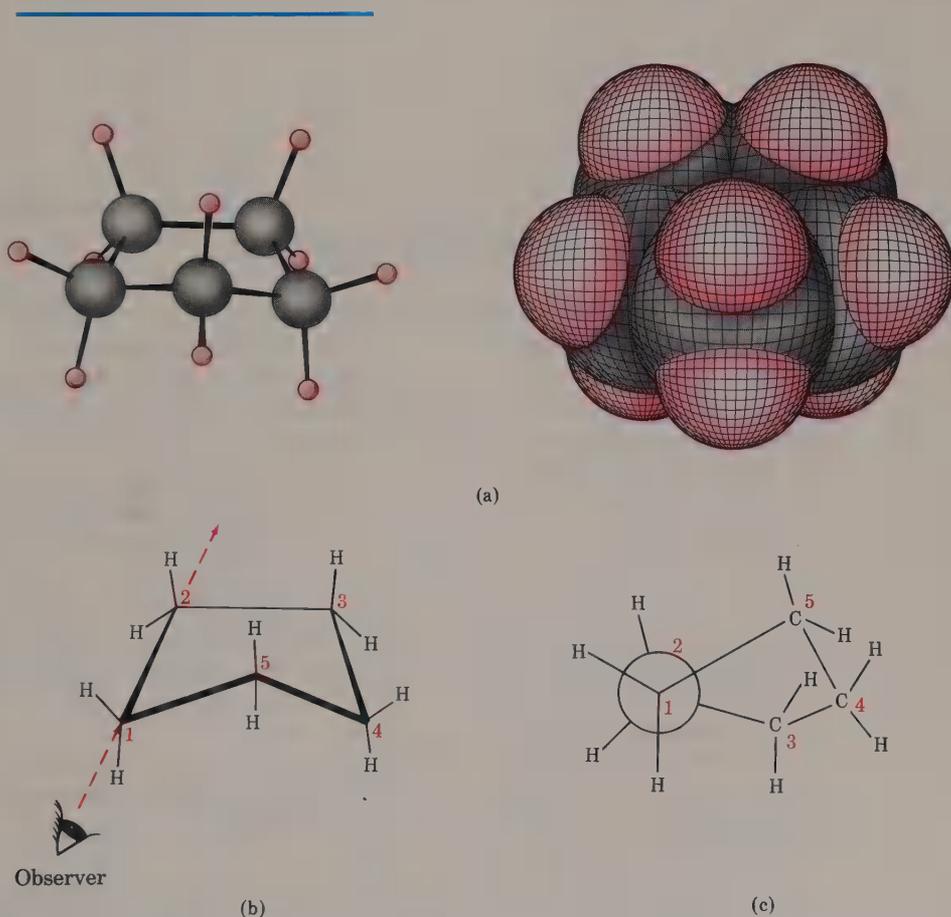
**Figure 4.10** An orbital view of bent bonds in cyclopropane: (a) normal alkane C-C bonds have head-on overlap, but (b) cyclopropane bent bonds overlap an angle. (c) An electron-density map provided by low-temperature X-ray studies shows how the electron density in the C3-C3' cyclopropane bond of the model compound is displaced off the internuclear axis, whereas the "normal" C3-C4 and C3'-C4' bonds are centered on their internuclear axes.



**Figure 4.11** The conformation of cyclobutane. Part (c) is a Newman projection along the C1-C2 bond showing that neighboring C-H bonds are not quite eclipsed.

one carbon atom lies about  $25^\circ$  above the plane of the other three (Figure 4.11). The effect of this slight bend is to *increase* angle strain but to *decrease* torsional strain, until a minimum-energy balance between the two opposing effects is achieved.

Cyclopentane was predicted by Baeyer to be nearly strain-free, but combustion measurements indicate that this is not the case. Although cyclopentane has practically no angle strain, the 10 pairs of neighboring hydrogens introduce considerable torsional strain into a planar conformation. As a result, cyclopentane adopts a puckered, out-of-plane conformation that strikes a balance between increased angle strain and decreased torsional strain. Four of the cyclopentane carbon atoms are in approximately the same plane, with the fifth carbon atom bent out of the plane. Most of the hydrogens are nearly staggered with respect to their neighbors (Figure 4.12).



**Figure 4.12** The conformation of cyclopentane. Carbons 2, 3, 4, and 5 are nearly planar, but carbon 1 is out of the plane. Part (c) is a Newman projection along the C1–C2 bond showing that neighboring C–H bonds are nearly staggered.

## PROBLEM .....

- 4.8 *cis*-1,2-Dimethylcyclobutane is less stable than its *trans* isomer, but *cis*-1,3-dimethylcyclobutane is more stable than its *trans* isomer. Explain.

## PROBLEM .....

- 4.9 Draw the most favored conformation of *cis*-1,3-dimethylcyclobutane (Problem 4.8).

## PROBLEM .....

- 4.10 How many hydrogen–hydrogen eclipsing interactions would be present if cyclopentane were planar? Assuming an energy cost of 1.0 kcal/mol for each eclipsing interaction, how much total strain would you expect planar cyclopentane to have? How much of this strain is relieved by puckering if the measured total strain of cyclopentane is 6.5 kcal/mol?
- .....

## 4.9 Conformations of Cyclohexane

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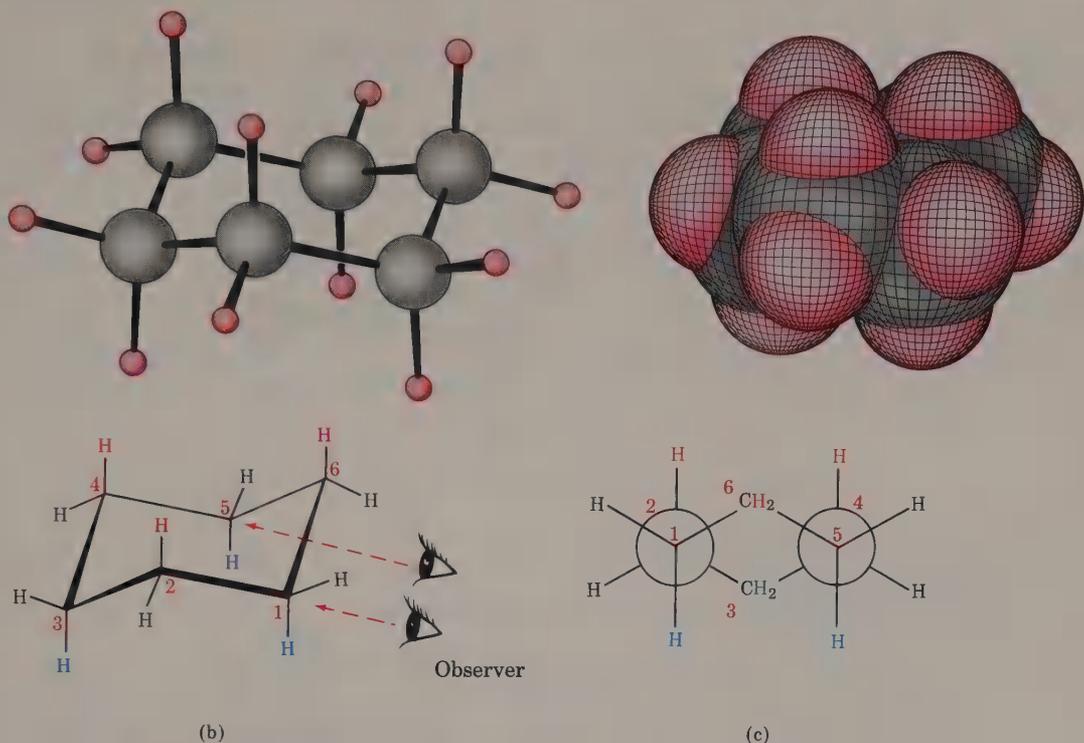
Substituted cyclohexanes are the most important of all cycloalkanes because of their wide occurrence in nature. A large number of compounds, including many important pharmaceutical agents, contain cyclohexane rings. Combustion data show that cyclohexane rings are strain-free, with neither angle strain nor torsional strain. How can this be?

The answer was first suggested in 1890 by H. Sachse and later expanded on by Ernst Mohr.<sup>3</sup> Cyclohexane rings are not flat as Baeyer assumed; they are puckered into a three-dimensional conformation that relieves all strain. The C–C–C angles of cyclohexane can reach the strain-free tetrahedral value if the ring adopts the **chair conformation** shown in Figure 4.13, so called because of its similarity to a lounge chair—a back, a seat, and a footrest. Furthermore, if we sight along any one of the carbon–carbon bonds in a Newman projection, we find that chair cyclohexane has no torsional strain; all neighboring C–H bonds are perfectly staggered (Figure 4.13).

The simplest way to visualize strain-free chair cyclohexane is to build your own molecular model. Two-dimensional drawings such as Figure 4.13 are useful, but there is no substitute for holding, twisting, and turning a three-dimensional model in your hands.

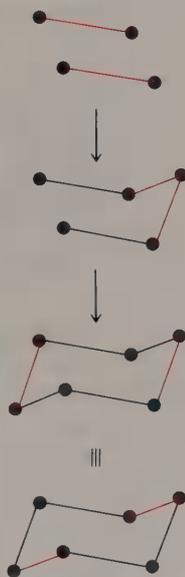
The chair conformation of cyclohexane is so important that all organic chemists must learn how to draw it properly by following the three steps in Figure 4.14.

<sup>3</sup>Ernst Mohr (1873–1926); b. Dresden; Ph.D. Kiel (1897); professor, University of Heidelberg.



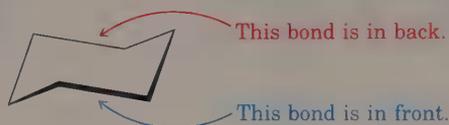
**Figure 4.13** The strain-free chair conformation of cyclohexane. All C–C–C bond angles are  $111.5^\circ$ , close to the ideal  $109.5^\circ$  tetrahedral angle.

1. Draw two parallel lines, slanted downward and slightly offset from each other. This means that four of the cyclohexane carbon atoms lie in a plane.
2. Locate the topmost carbon atom above and to the right of the plane of the other four and connect the bonds.
3. Locate the bottommost carbon atom below and to the left of the plane of the middle four and connect the bonds. Note that the bonds to the bottommost carbon atom are parallel to the bonds to the topmost carbon.



**Figure 4.14** How to draw the cyclohexane chair conformation.

It's important to remember when viewing chair cyclohexane that the lower bond is in front and the upper bond is in back. If this convention isn't defined, an optical illusion can make it appear that the reverse is true. For clarity, the cyclohexane rings drawn in this book will have the front (lower) bond heavily shaded to indicate its nearness to the viewer.



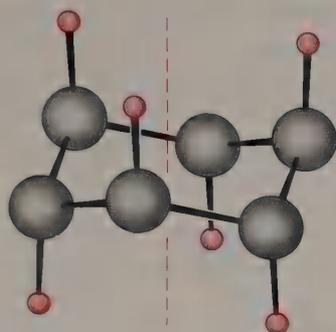
## 4.10 Axial and Equatorial Bonds in Cyclohexane

The chair conformation of cyclohexane has many consequences. For example, we'll see in Section 11.12 that the chemical behavior of many substituted cyclohexanes is directly controlled by their conformation. Another consequence of the chair cyclohexane conformation is that there are two kinds of hydrogen atoms on the ring: **axial hydrogens** and **equatorial hydrogens** (Figure 4.15). Chair cyclohexane has six axial hydrogens that are perpendicular to the ring (parallel to the ring *axis*) and six equatorial hydrogens that are more or less in the rough plane of the ring (around the ring *equator*).

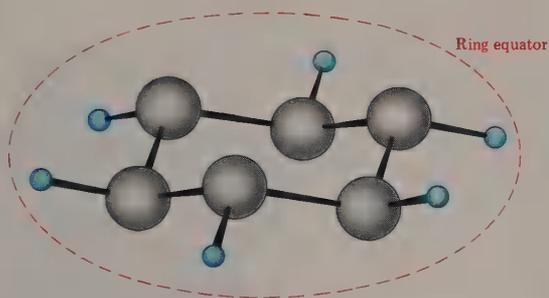
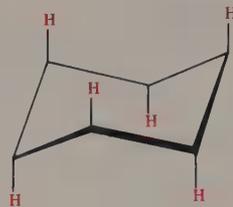
Look carefully at the disposition of the axial and equatorial hydrogens in Figure 4.15. Each carbon atom in cyclohexane has one axial and one equatorial hydrogen. Furthermore, every cyclohexane ring has two sides, or "faces." Each face has both axial and equatorial hydrogens in an alternating axial–equatorial–axial–equatorial arrangement. For example, the top face of the cyclohexane ring shown in Figure 4.16 has axial hydrogens on carbons 1, 3, and 5, but has equatorial hydrogens on carbons 2, 4, and 6. Exactly the reverse is true for the bottom face: Carbons 1, 3, and 5 have equatorial hydrogens, but carbons 2, 4, and 6 have axial hydrogens.

Note that we haven't used the words *cis* and *trans* in this discussion of cyclohexane geometry. Two hydrogens on the same face of the ring are always *cis*, regardless of whether they're axial or equatorial and regardless of whether they're adjacent. Similarly, two hydrogens on opposite faces of the ring are always *trans*, regardless of whether they're axial or equatorial. Axial and equatorial bonds can be drawn in the following way (look at a molecular model as you practice):

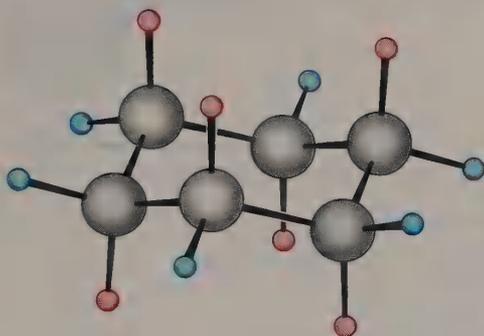
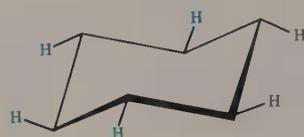
1. *Axial bonds* The six axial bonds, one on each carbon, are parallel and have an alternating up–down relationship.
2. *Equatorial bonds* The six equatorial bonds come in three sets of two parallel lines. Each set is also parallel to two ring bonds, as shown in blue on the structural diagram on page 122.



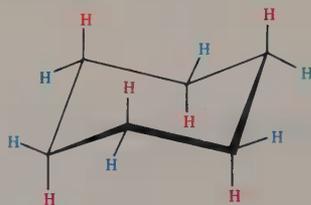
(a) Six *axial* hydrogens (parallel to axis of ring)



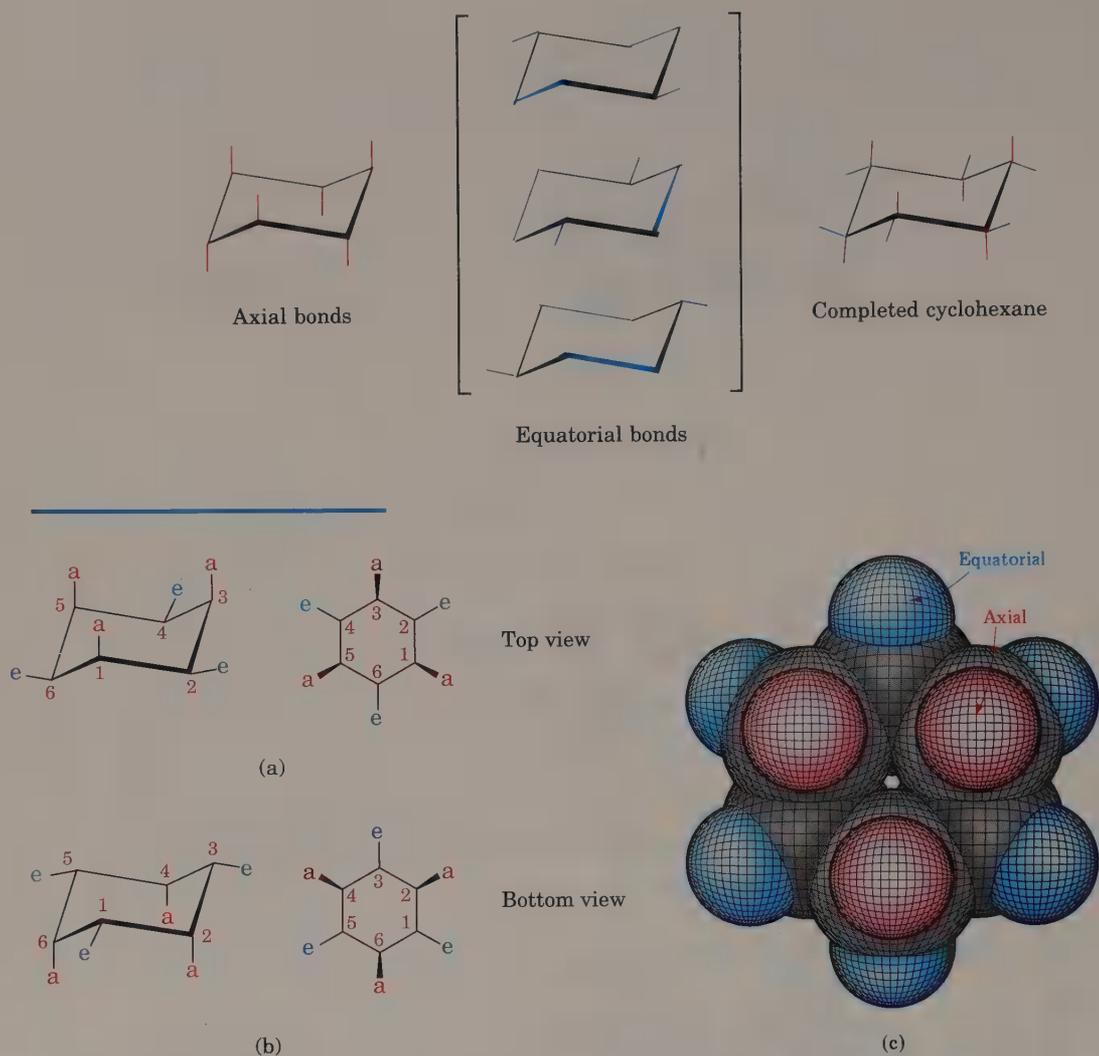
(b) Six *equatorial* hydrogens (in a band around the equator of the ring)



(c) Chair cyclohexane with all its hydrogen atoms



**Figure 4.15** Axial and equatorial hydrogen atoms in chair cyclohexane.

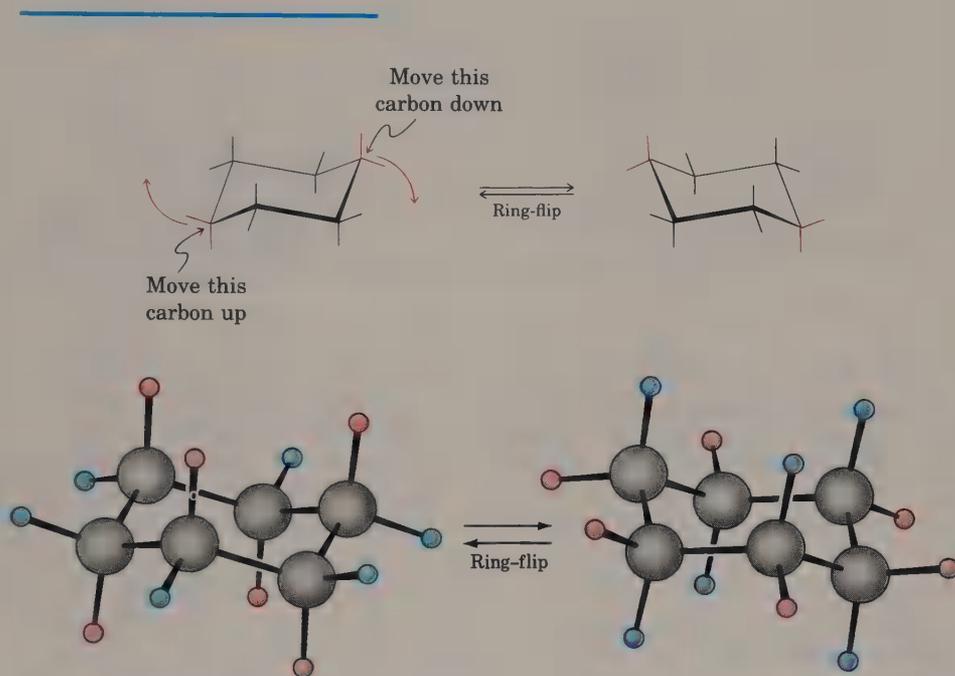


**Figure 4.16** Alternating axial and equatorial positions in chair cyclohexane. Each carbon atom has both an axial and an equatorial position, and each face has alternating axial/equatorial positions. Part (c) is a computer-generated picture of a cyclohexane ring viewed from the top.

## 4.11 Conformational Mobility of Cyclohexane

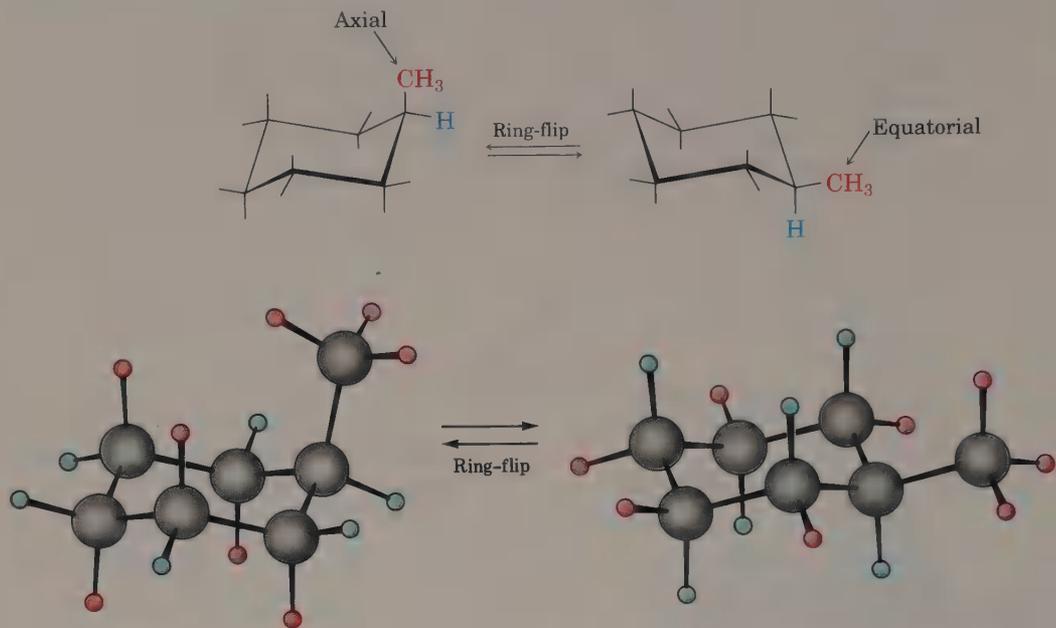
Since chair cyclohexane has two kinds of positions, axial and equatorial, we might expect to find two isomeric forms of a monosubstituted cyclohexane. In fact, though, this expectation is wrong. At room temperature, there is

only *one* methylcyclohexane, *one* bromocyclohexane, *one* cyclohexanol, and so on because cyclohexane rings are *conformationally mobile*. The two chair conformations can readily interconvert, with the result that axial and equatorial positions become interchanged. The interconversion of chair conformations, usually referred to as a **ring-flip**, is shown in Figure 4.17. Molecular models show the process more clearly, and you should practice ring-flipping with models.



**Figure 4.17** A ring-flip in chair cyclohexane interconverts axial and equatorial positions.

We can mentally ring-flip a chair cyclohexane by holding the middle four carbon atoms in place while folding the two ends in opposite directions. The net result of carrying out a ring-flip is that axial and equatorial positions become interconverted. An axial substituent in one chair form becomes an equatorial substituent in the ring-flipped chair form, and vice versa. For example, axial methylcyclohexane becomes equatorial methylcyclohexane after ring-flip (Figure 4.18). Spectroscopic measurements indicate that the energy barrier to chair–chair interconversions is about 10.8 kcal/mol (45 kJ/mol), a value low enough to make the process extremely rapid at room temperature. We therefore see only a single, rapidly interconverting average structure, rather than distinct axial and equatorial isomers.

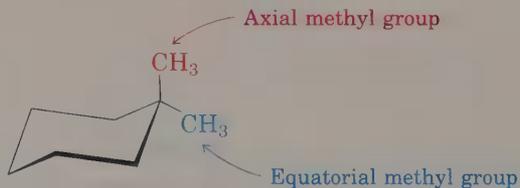


**Figure 4.18** A ring-flip converts axial methylcyclohexane into equatorial methylcyclohexane.

**PRACTICE PROBLEM**.....

Draw 1,1-dimethylcyclohexane, indicating whether each methyl group is axial or equatorial.

**Solution** First draw a chair cyclohexane ring and then put two methyl groups on the same carbon. The methyl group in the rough plane of the ring must be equatorial, and the other (above or below the ring) must be axial.



**PROBLEM**.....

- 4.11 Draw two different chair conformations of bromocyclohexane, showing all hydrogen atoms. Identify each substituent as axial or equatorial.

## PROBLEM.....

- 4.12 Explain why a *cis*-1,2-disubstituted cyclohexane such as *cis*-1,2-dichlorocyclohexane must have one group axial and one group equatorial.

## PROBLEM.....

- 4.13 Explain why a *trans*-1,2-disubstituted cyclohexane must have either both groups axial or both groups equatorial.

## PROBLEM.....

- 4.14 Draw two different chair conformations of *trans*-1,4-dimethylcyclohexane, and label all positions as axial or equatorial.
- .....

## 4.12 Conformations of Monosubstituted Cyclohexanes

---

Let's look at the consequences of axial and equatorial cyclohexane bonds in substituted cyclohexanes. Although cyclohexane rings rapidly flip conformations at room temperature, the two conformers of a monosubstituted cyclohexane aren't equally stable. In methylcyclohexane, for example, the equatorial conformer is more stable than the axial conformer by 1.8 kcal/mol (8 kJ/mol). Similarly for other monosubstituted cyclohexanes: A substituent is almost always more stable in an equatorial position than in an axial position.

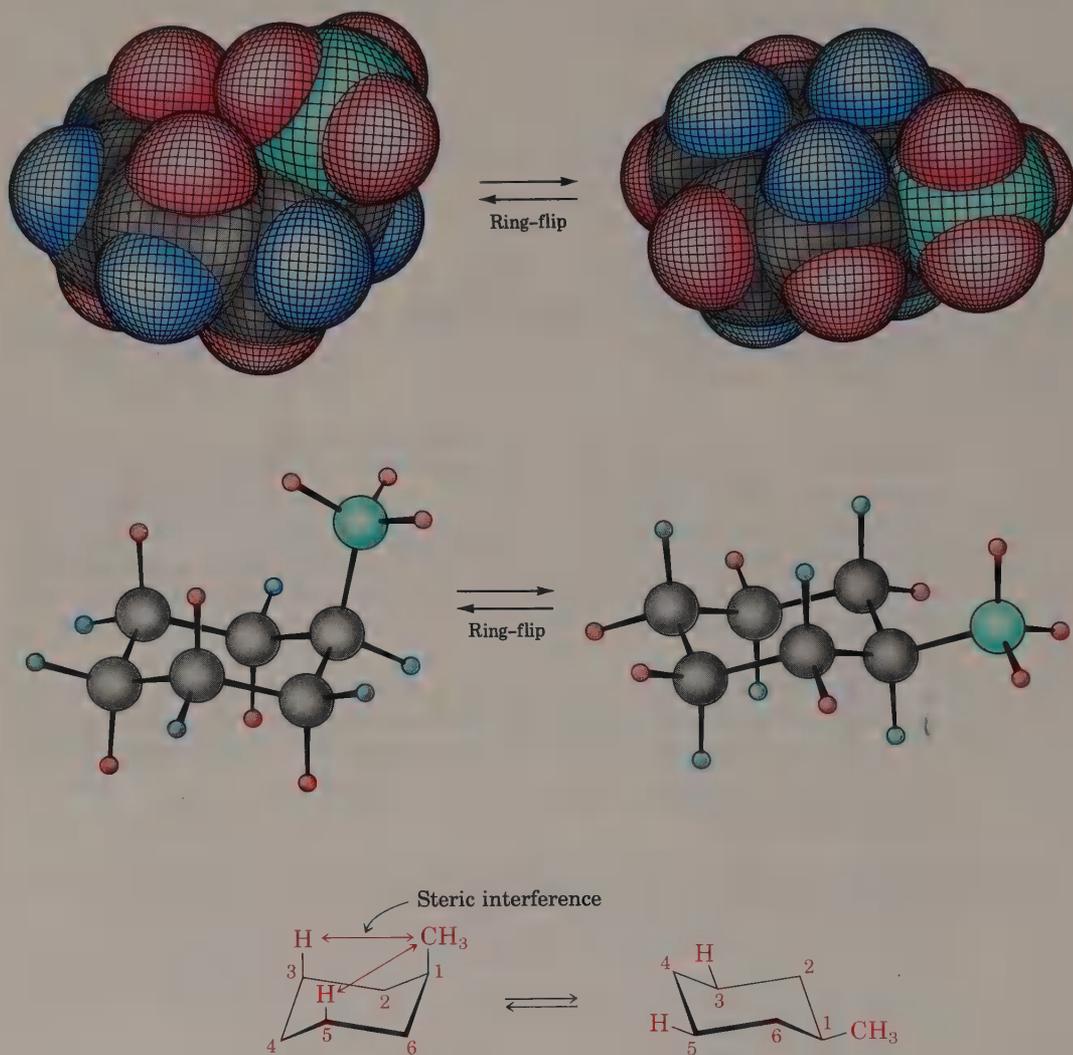
It's possible to calculate the exact percentages of any two isomers at equilibrium, as shown in Table 4.3. For example, an energy difference of 1.8 kcal/mol means that about 95% of methylcyclohexane molecules have the methyl group equatorial at any given instant, and only 5% have the methyl group axial.

**TABLE 4.3** The Relationship Between Stability and Isomer Percentages at Equilibrium<sup>a</sup>

More stable isomer (%)	Less stable isomer (%)	Energy difference at 25°C	
		(kcal/mol)	(kJ/mol)
50	50	0	0
75	25	0.651	2.72
90	10	1.302	5.45
95	5	1.744	7.29
99	1	2.722	11.38
99.9	0.1	4.092	17.11

<sup>a</sup>The values in this table are calculated from the equation  $K = e^{-(\Delta E/RT)}$ , where  $K$  is the equilibrium constant between isomers;  $e \approx 2.718$  (the base of natural logarithms);  $\Delta E$  = energy difference between isomers;  $T$  = absolute temperature (in kelvins); and  $R = 1.986 \text{ cal/mol} \cdot K$  (the gas constant).

The energy difference between axial and equatorial conformers is due to steric strain caused by so-called **1,3-diaxial interactions**. That is, the axial methyl group on C1 is too close to the axial hydrogens three carbons away on C3 and C5, resulting in 1.8 kcal/mol of steric strain (Figure 4.19).



**Figure 4.19** Interconversion of axial and equatorial methylcyclohexane. The equatorial conformation is more stable by 1.8 kcal/mol.

1,3-Diaxial steric strain is already familiar; we've seen it before as the steric strain between methyl groups in gauche butane (Section 4.3). Recall that gauche butane is less stable than anti butane by 0.9 kcal/mol because of steric interference between hydrogen atoms on the two methyl groups. If

we look at a four-carbon fragment of axial methylcyclohexane as a gauche butane, it's apparent that the steric interaction is the same in both cases. Since methylcyclohexane has two such interactions, however, there is  $2 \times 0.9 = 1.8$  kcal/mol steric strain. The origin of 1,3-diaxial steric strain is shown in Figure 4.20.

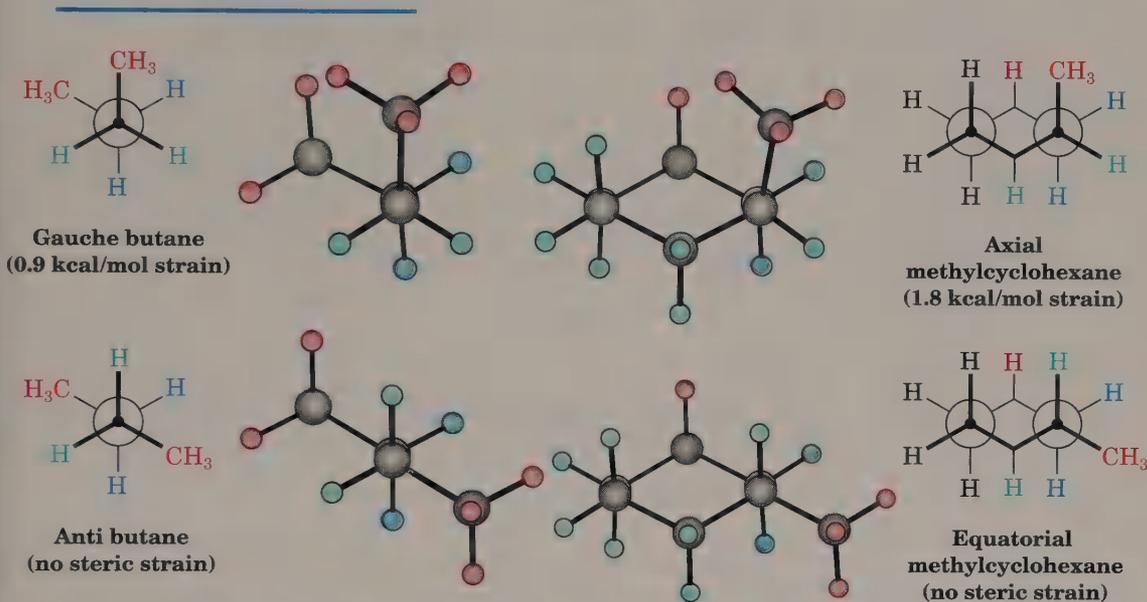
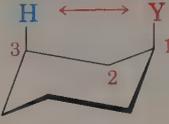


Figure 4.20 Origin of 1,3-diaxial cyclohexane interactions.

Sighting along the C1–C2 bond of axial methylcyclohexane shows that the axial hydrogen at C3 has a gauche butane interaction with the axial methyl group at C1. Sighting similarly along the C1–C6 bond shows that the axial hydrogen at C5 has a gauche butane interaction with the axial methyl group at C1. Both of these interactions are absent in equatorial methylcyclohexane, and we therefore find an energy difference of 1.8 kcal/mol (8 kJ/mol) between the two forms.

What's true for methylcyclohexane is also true for other monosubstituted cyclohexanes: A substituent is almost always more stable in an equatorial position than in an axial position. The exact amount of 1,3-diaxial steric strain in a specific compound depends, of course, on the nature and size of the axial group. Table 4.4 lists strain values for some common substituents. As you might expect, the amount of steric strain increases through the series  $\text{H}_3\text{C}- < \text{CH}_3\text{CH}_2- < (\text{CH}_3)_2\text{CH}- \ll (\text{CH}_3)_3\text{C}-$  in parallel with the increasing bulk of the successively larger alkyl groups. Note that the values in Table 4.4 refer to 1,3-diaxial interactions of the indicated group with a *single* hydrogen atom. These values must therefore be doubled to arrive at the amount of strain in a monosubstituted cyclohexane.

TABLE 4.4 Steric Strain Due to 1,3-Diaxial Interactions

Y	Strain of one H-Y 1,3-diaxial interaction		
	(kcal/mol)	(kJ/mol)	
—F	0.12	0.5	
—Cl	0.25	1.4	
—Br	0.25	1.4	
—OH	0.5	2.1	
—CH <sub>3</sub>	0.9	3.8	
—CH <sub>2</sub> CH <sub>3</sub>	0.95	4.0	
—CH(CH <sub>3</sub> ) <sub>2</sub>	1.1	4.6	
—C(CH <sub>3</sub> ) <sub>3</sub>	2.7	11.3	
—C <sub>6</sub> H <sub>5</sub>	1.5	6.3	
—COOH	0.7	2.9	
—CN	0.1	0.4	

## PROBLEM .....

- 4.15 How can you account for the fact (Table 4.4) that an axial *tert*-butyl substituent has much larger 1,3-diaxial interactions than isopropyl, but isopropyl is fairly similar to ethyl and methyl? Use molecular models to help with your answer.

## PROBLEM .....

- 4.16 Why do you suppose an axial cyano substituent causes practically no 1,3-diaxial steric strain (0.1 kcal/mol)?

## PROBLEM .....

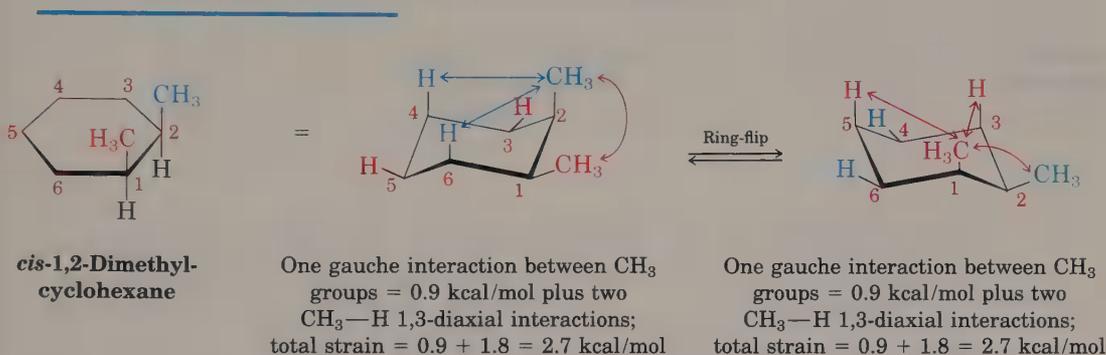
- 4.17 Look at Table 4.3 and calculate the percentages of axial and equatorial conformers present at equilibrium in *tert*-butylcyclohexane.
- .....

## 4.13 Conformational Analysis of Disubstituted Cyclohexanes

Monosubstituted cyclohexanes usually have the substituent equatorial. In disubstituted cyclohexanes, however, the situation is more complex because the steric effects of both substituents must be taken into account. All steric

interactions in both possible conformations must be analyzed before deciding which conformation is favored.

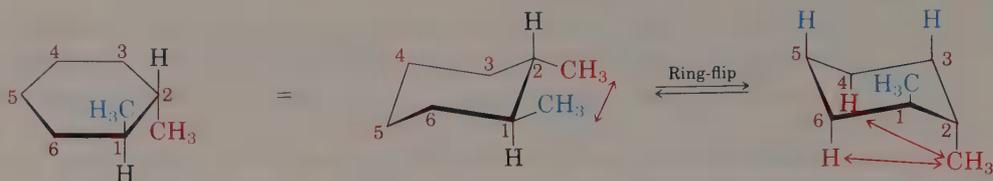
Let's look first at 1,2-dimethylcyclohexane. There are two isomers, *cis*-1,2-dimethylcyclohexane and *trans*-1,2-dimethylcyclohexane, and we have to consider them separately. In the *cis* isomer, both methyl groups are on the same side of the ring, and the compound can exist in either of the two chair conformations shown in Figure 4.21. (Note that it's often easier to see whether a compound is *cis*- or *trans*-disubstituted by first drawing a flat representation and then converting to a chair conformation.)



**Figure 4.21** Conformations of *cis*-1,2-dimethylcyclohexane. The two conformations are equal in energy since each has one axial methyl group and one equatorial methyl group.

Both conformations of *cis*-1,2-dimethylcyclohexane in Figure 4.21 have one methyl group axial and one methyl group equatorial. The conformation on the left has an axial methyl group at C2 that has 1,3-diaxial interactions with hydrogens on C4 and C6. The ring-flipped conformation on the right has an axial methyl group at C1 that has 1,3-diaxial interactions with hydrogens on C3 and C5. In addition, both conformations have gauche butane interactions between the two methyl groups. The *two conformations are exactly equal in energy*, with a total steric strain of 2.7 kcal/mol (11.3 kJ/mol).

In *trans*-1,2-dimethylcyclohexane, the two methyl groups are on opposite sides of the ring, and the compound can exist in either of the two chair conformations shown in Figure 4.22. The situation here is quite different from that of the *cis* isomer. The *trans* conformation on the left in Figure 4.22 has both methyl groups equatorial, and therefore it has only a gauche butane interaction between methyls and no 1,3-diaxial interactions. The conformation on the right, however, has both methyl groups axial. The axial methyl group at C1 interacts with axial hydrogens at C3 and C5, and the axial methyl group at C2 interacts with axial hydrogens at C4 and C6. These four 1,3-diaxial interactions make the diaxial conformation  $3.6 - 0.9 = 2.7$  kcal/mol less favorable than the diequatorial conformation. We therefore predict that *trans*-1,2-dimethylcyclohexane will exist almost exclusively (> 99%) in the diequatorial conformation.



*trans*-1,2-Dimethylcyclohexane One gauche interaction between  $\text{CH}_3$  groups = 0.9 kcal/mol

Four 1,3-diaxial interactions = 3.6 kcal/mol

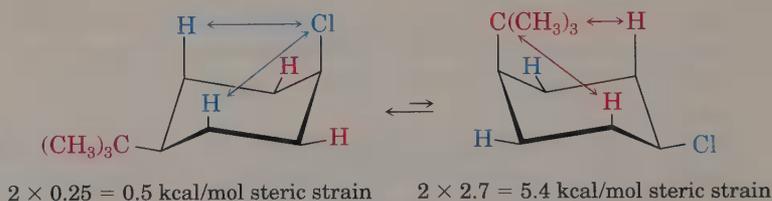
**Figure 4.22** Conformations of *trans*-1,2-dimethylcyclohexane. The conformation with both methyl groups equatorial is favored by 2.7 kcal/mol.

The same kind of **conformational analysis** just carried out for 1,2-dimethylcyclohexane can be carried out for any substituted cyclohexane. For example, let's look at *cis*-1-*tert*-butyl-4-chlorocyclohexane in the following practice problem. It turns out that the large amount of steric strain caused by an axial *tert*-butyl group effectively holds the cyclohexane ring in a single conformation. We can sometimes take advantage of this steric locking if we want to study the chemical reactivity of an immobile cyclohexane ring.

#### PRACTICE PROBLEM.....

What is the most stable conformation of *cis*-1-*tert*-butyl-4-chlorocyclohexane, and by how much is it favored?

**Solution** First draw the two chair conformations of the molecule. *cis*-1-*tert*-Butyl-4-chlorocyclohexane can exist in either of the two chair conformations indicated:



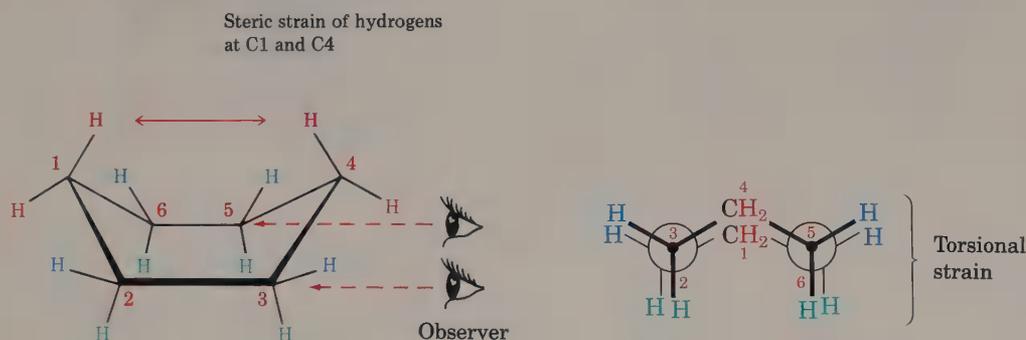
In the left-hand conformation, the *tert*-butyl group is equatorial and the chlorine is axial. In the right-hand conformation, the *tert*-butyl group is axial and the chlorine is equatorial. These conformations aren't of equal energy because an axial *tert*-butyl substituent and an axial chloro substituent produce different amounts of steric strain. Table 4.4 shows that a single *tert*-butyl-hydrogen 1,3-diaxial interaction costs 2.7 kcal/mol, whereas a single chlorine-hydrogen 1,3-diaxial interaction costs only 0.25 kcal/mol. An axial *tert*-butyl group therefore induces  $(2 \times 2.7) - (2 \times 0.25) = 4.9$  kcal/mol more steric strain than does an axial chlorine, and the compound adopts the conformation having the chlorine axial and the *tert*-butyl equatorial.

## PROBLEM .....

- 4.18 Draw the most stable chair conformation of these molecules, and estimate the amount of 1,3-diaxial strain in each.
- (a) *trans*-1-Chloro-3-methylcyclohexane      (b) *cis*-1-Ethyl-2-methylcyclohexane  
 (c) *cis*-1-Bromo-4-ethylcyclohexane      (d) *cis*-1-*tert*-Butyl-4-ethylcyclohexane
- .....

## 4.14 Boat Cyclohexane

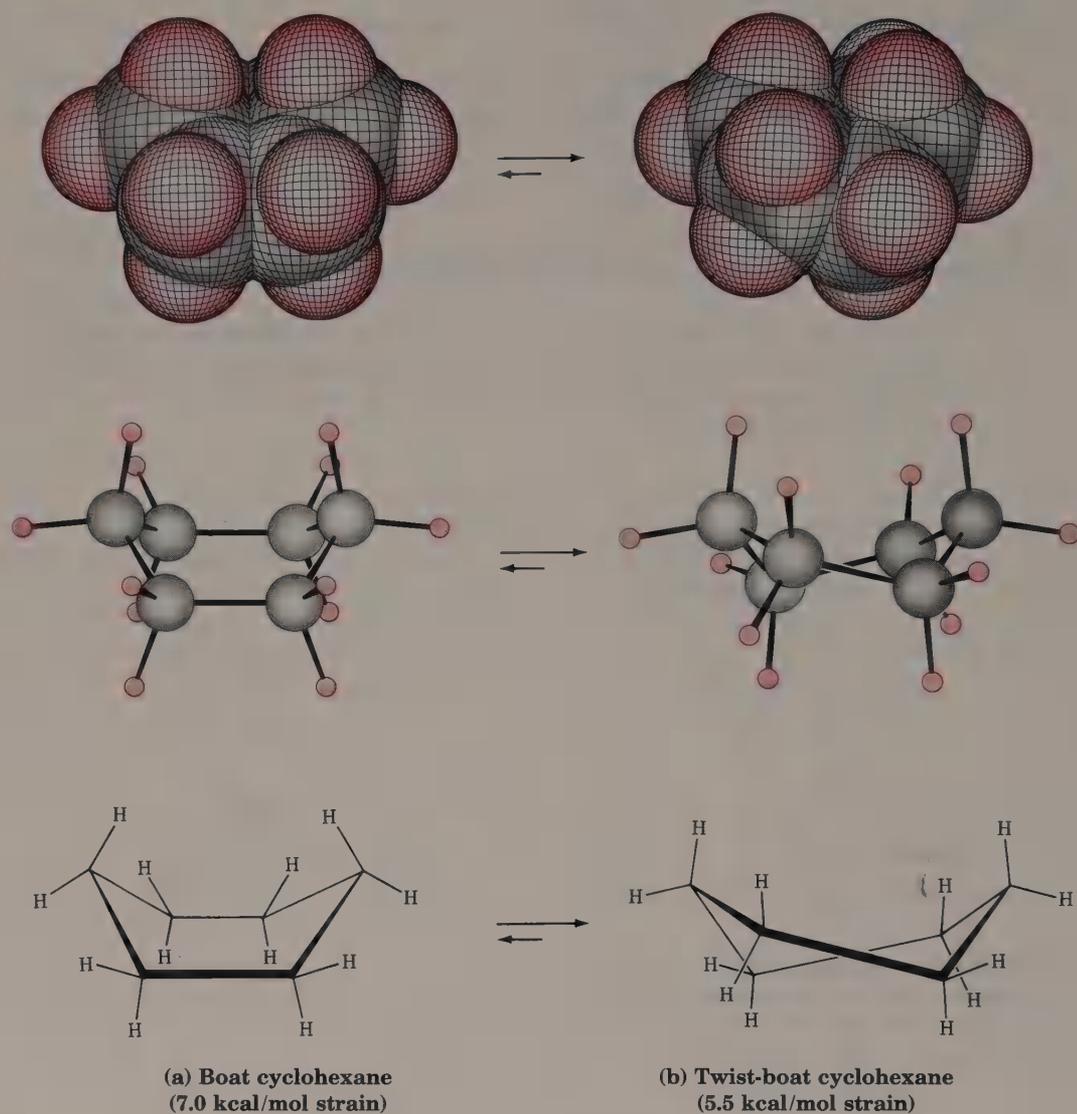
In addition to chair cyclohexane, a second conformation known as **boat cyclohexane** is also free of angle strain. We haven't paid it any attention thus far, however, because boat cyclohexane is much less stable than chair cyclohexane (Figure 4.23).



**Figure 4.23** The boat conformation of cyclohexane. There is no angle strain in this conformation, but there is both steric strain and torsional strain.

There are two kinds of carbon atoms in boat cyclohexane. Carbons 2, 3, 5, and 6 lie in a plane, with carbons 1 and 4 above the plane. The inside hydrogen atoms on carbons 1 and 4 approach each other closely enough to cause considerable steric strain, and the four pairs of hydrogens on carbons 2, 3, 5, and 6 are eclipsed. Thus, boat cyclohexane has both steric strain and torsional strain. The Newman projection in Figure 4.23, obtained by sighting along the C2–C3 and C5–C6 bonds, shows this eclipsing clearly.

Measurements indicate that boat cyclohexane is approximately 7.0 kcal/mol (29 kJ/mol) less stable than chair cyclohexane, although this value is reduced to about 5.5 kcal/mol by twisting slightly, thereby relieving some torsional strain (Figure 4.24). Even the **twist-boat conformation** is still much more strained than the chair conformation, though, and molecules adopt this geometry only under special circumstances.



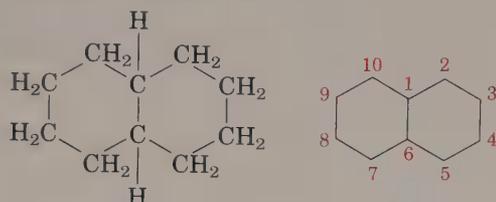
**Figure 4.24** Boat and twist-boat conformations of cyclohexane. The twist-boat conformation is lower in energy than the boat conformation by 1.5 kcal/mol. Both conformations are much more strained than chair cyclohexane.

**PROBLEM** .....

- 4.19 *trans*-1,3-di-*tert*-Butylcyclohexane exists largely in a twist-boat conformation. Why might this be? Draw the likely twist-boat conformation.
- .....

## 4.15 Conformations of Polycyclic Molecules

The last point we'll consider about cycloalkane stereochemistry is what happens when two or more cycloalkane rings are fused together to construct a **polycyclic** molecule such as decalin.



Decalin (two fused cyclohexane rings)

Decalin consists of two cyclohexane rings joined together to share two carbon atoms (the **bridgehead** carbons, C1 and C6) and a common bond. Decalin can exist in either of two isomeric forms, depending on whether the rings are *trans* fused or *cis* fused. In *trans*-decalin, the hydrogen atoms at the bridgehead carbons are on opposite sides of the rings; in *cis*-decalin, the bridgehead hydrogens are on the same side. Figure 4.25 shows how both compounds can be represented using chair cyclohexane conformations.

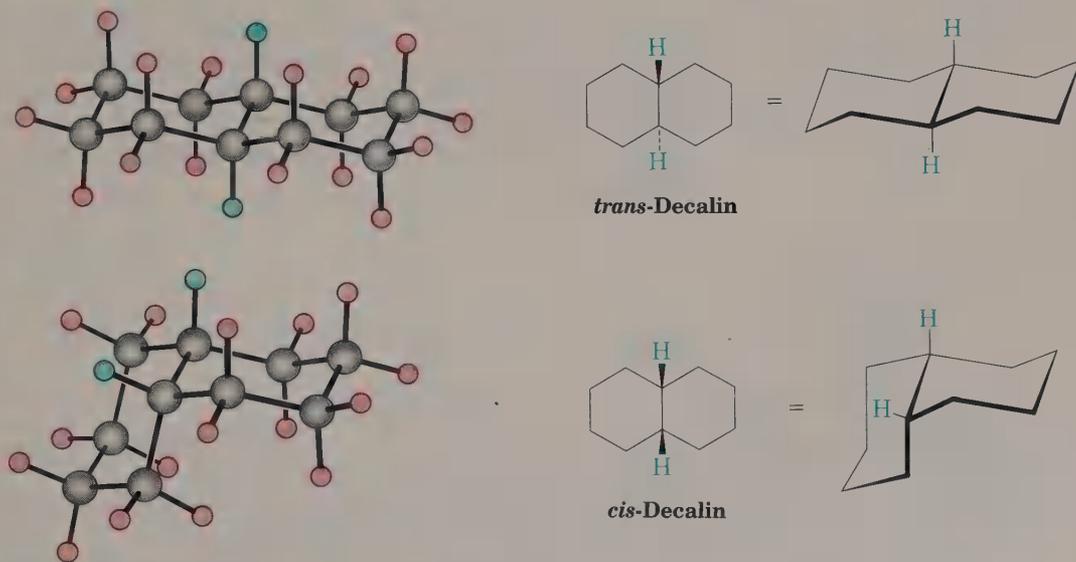
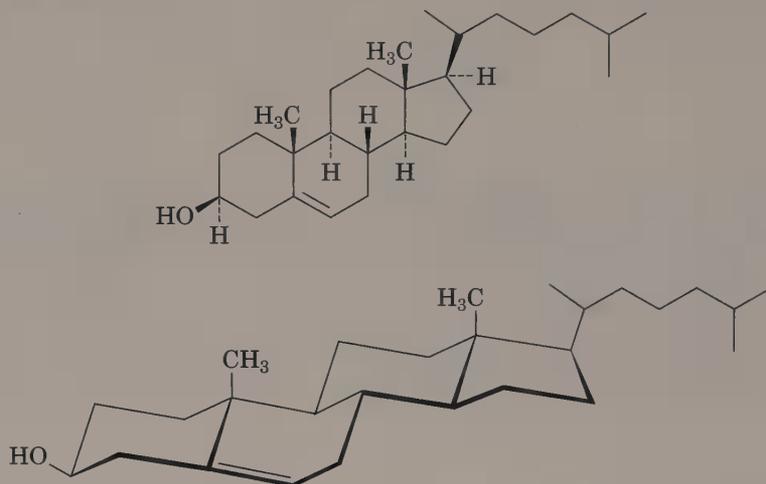


Figure 4.25 Representations of *trans*- and *cis*-decalin.

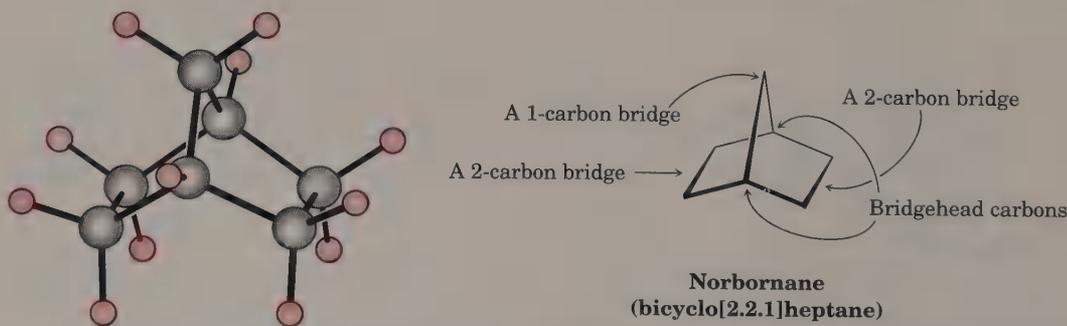
Note that *trans*- and *cis*-decalin are not interconvertible by ring-flips or other rotations. They are stereoisomers (Section 3.9) and bear the same relationship to each other that *cis*- and *trans*-1,2-dimethylcyclohexane do.

Polycyclic compounds are extremely important, and many valuable substances have fused-ring structures. For example, the steroids, such as cholesterol, have four rings—three six-membered and one five-membered—fused together. Though steroids look complicated compared with cyclohexane or decalin, they consist simply of chair cyclohexane rings locked together. The same principles that apply to the conformational analysis of simple cyclohexane rings apply equally well (and often better) to steroids.



Cholesterol (a steroid)

Another common fused-ring system is the norbornane or bicyclo[2.2.1]heptane structure. Norbornane is a *bicycloalkane*, so called because two rings would have to be broken open in order to generate an acyclic structure. Its systematic name, bicyclo[2.2.1]heptane, reflects the fact that it has a total of seven carbons, is bicyclic, and has three “bridges” of 2, 2, and 1 carbon atoms connecting the two “bridgehead” carbons.

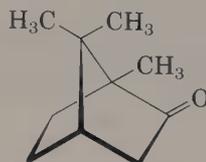


Norbornane  
(bicyclo[2.2.1]heptane)

Norbornane has a conformationally locked boat cyclohexane ring in which carbons 1 and 4 are joined by an extra  $-\text{CH}_2-$  group. Note how, in

drawing this structure, we show one bond crossing in front of another by indicating a break in the rear bond. Making a molecular model is particularly helpful when trying to see the three-dimensionality of norbornane.

Substituted norbornanes, such as camphor, are found widely in nature, and many have played important historical roles in developing organic structural theories.



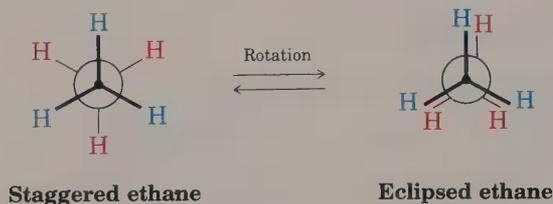
Camphor

PROBLEM.....

- 4.20 Which isomer is more stable, *trans*-decalin or *cis*-decalin? Explain your answer in terms of the steric interactions present in the two molecules.
- .....

## 4.16 Summary and Key Words

Carbon-carbon single bonds in alkanes are formed by sigma overlap of two carbon  $sp^3$  hybrid orbitals. Rotation is possible around sigma bonds because of their cylindrical symmetry, and alkanes therefore have a large number of rapidly interconverting **conformations**. **Newman projections** make it possible to visualize the spatial consequences of bond rotation by sighting directly along a carbon-carbon bond axis. The **staggered** conformation of ethane is 2.9 kcal/mol (12 kJ/mol) more stable than the **eclipsed** conformation because of **torsional strain**. In general, any alkane is most stable when all of its bonds are staggered.



Most organic molecules contain saturated rings. These alicyclic (aliphatic cyclic) molecules have special characteristics that affect the chemistry they undergo. Combustion studies indicate that not all cycloalkanes are equally stable. Cyclopropane and cyclobutane are the least stable rings, whereas cyclopentane and the medium-sized rings (cycloheptane through cyclotridecane) all have varying degrees of strain. Cyclohexane and large rings (cyclotetradecane and above) are virtually strain-free.

Three different kinds of strain contribute to the overall energy level of a cycloalkane: (1) **angle strain**, the resistance of a bond angle to compression or expansion from the normal  $109^\circ$  tetrahedral value; (2) **torsional strain**, the energy cost of having neighboring C–H bonds eclipsed rather than staggered; and (3) **steric strain**, the result of the repulsive van der Waals interactions that arise when two groups try to occupy the same spatial position.

Cyclopropanes are highly strained (27.6 kcal/mol) because of both angle strain and torsional strain. Cyclobutane is also highly strained (26.4 kcal/mol) because of both angle strain and torsional strain. Cyclopentane is free of angle strain but has a large number of eclipsing interactions. Both cyclobutane and cyclopentane pucker slightly away from planarity in order to relieve torsional strain.

Cyclohexane rings are the most important of all ring sizes because of their wide occurrence. Cyclohexane is strain-free because of its puckered **chair conformation**, in which all bond angles are near  $109^\circ$  and all neighboring C–H bonds are staggered. Chair cyclohexane has two kinds of hydrogens: **axial** and **equatorial**. Axial hydrogens are directed up and down, parallel to the ring axis, whereas equatorial hydrogens lie in a belt more or less along the equator of the ring. Each carbon atom has one axial and one equatorial hydrogen.

Chair cyclohexanes are conformationally mobile and can undergo a **ring-flip** that interconverts axial and equatorial positions:



Substituents on the ring are more stable in the equatorial position, since axial substituents cause **1,3-diaxial steric strain**. The amount of 1,3-diaxial strain caused by an axial substituent depends on its bulk. The stereochemistry of cyclohexane and its derivatives is best learned by using molecular models to examine conformational relationships.

## ADDITIONAL PROBLEMS .....

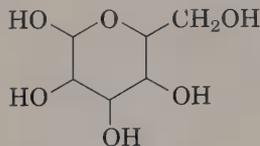
- 4.21 Define these terms in your own words:
- |                        |                   |                      |
|------------------------|-------------------|----------------------|
| (a) Angle strain       | (b) Steric strain | (c) Torsional strain |
| (d) Heat of combustion | (e) Conformation  | (f) Staggered        |
| (g) Eclipsed           | (h) Gauche butane |                      |
- 4.22 Consider 2-methylbutane (isopentane). Sighting along the C2–C3 bond:
- Draw a Newman projection of the most stable conformation.
  - Draw a Newman projection of the least stable conformation.
  - If a  $\text{CH}_3\text{--CH}_3$  eclipsing interaction costs 2.5 kcal/mol and a  $\text{CH}_3\text{--CH}_3$  gauche interaction costs 0.9 kcal/mol, construct a quantitative diagram of energy versus rotation about the C2–C3 bond.
- 4.23 What are the energy differences between the three possible staggered conformations around the C2–C3 bond in 2,3-dimethylbutane?

- 4.24 Construct a qualitative potential-energy diagram for rotation about the C–C bond of 1,2-dibromoethane. Which conformation would you expect to be more stable? Label the anti and the gauche conformations of 1,2-dibromoethane.
- 4.25 Which conformation of 1,2-dibromoethane (Problem 4.24) would you expect to have the larger dipole moment? The observed dipole moment is  $\mu = 1.0$  D. What does this tell you about the actual structure of the molecule?
- 4.26 The barrier to rotation about the C–C bond in bromoethane is 3.6 kcal/mol (15 kJ/mol).
- What energy value can you assign to an H–Br eclipsing interaction?
  - Construct a quantitative diagram of potential energy versus amount of bond rotation for bromoethane.
- 4.27 Define these terms:
- Axial bond
  - Equatorial bond
  - Chair conformation
  - 1,3-Diaxial interaction
- 4.28 Draw a chair cyclohexane ring and label all positions as axial or equatorial.
- 4.29 Why is a 1,3-cis disubstituted cyclohexane always more stable than its trans isomer?
- 4.30 Why is a 1,2-trans disubstituted cyclohexane always more stable than its cis isomer?
- 4.31 Which is more stable, a 1,4-trans disubstituted cyclohexane or its cis isomer?
- 4.32 *N*-Methylpiperidine has the conformation shown. What does this tell you about the relative steric requirements of a methyl group versus an electron lone pair?



*N*-Methylpiperidine

- 4.33 Draw the two chair conformations of *cis*-1-chloro-2-methylcyclohexane. Which is more stable and by how much?
- 4.34 Draw the two chair conformations of *trans*-1-chloro-2-methylcyclohexane. Which is more stable and by how much?
- 4.35  $\beta$ -Glucose contains a six-membered ring in which all of the substituents are equatorial. Draw  $\beta$ -glucose in its more stable chair conformation.



Glucose

- 4.36 From the data in Tables 4.3 and 4.4, calculate the percentages of molecules that have their substituents in an axial orientation for these compounds:
- Isopropylcyclohexane
  - Bromocyclohexane
  - Cyclohexanecarbonitrile,  $C_6H_{11}CN$
  - Cyclohexanol,  $C_6H_{11}OH$
- 4.37 Assume that you have a variety of cyclohexanes substituted in the positions indicated. Identify the substituents as either axial or equatorial. For example, a 1,2-cis

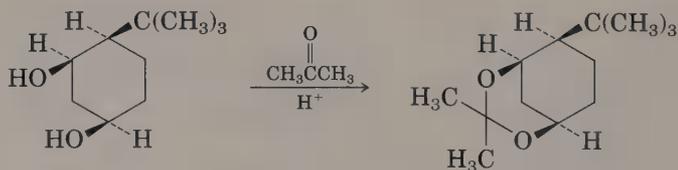
relationship means that one substituent must be axial and one equatorial, whereas a 1,2-*trans* relationship means that both substituents are axial or both are equatorial.

- (a) 1,3-*Trans* disubstituted                      (b) 1,4-*Cis* disubstituted  
 (c) 1,3-*Cis* disubstituted                        (d) 1,5-*Trans* disubstituted  
 (e) 1,5-*Cis* disubstituted                        (f) 1,6-*Trans* disubstituted
- 4.38 The diaxial conformation of *cis*-1,3-dimethylcyclohexane is approximately 5.4 kcal/mol (23 kJ/mol) less stable than the diequatorial conformation. Draw the two possible chair conformations and suggest a reason for the large energy difference.
- 4.39 Approximately how much steric strain does the 1,3-diaxial interaction between the two methyl groups (Problem 4.38) introduce into the diaxial conformation?
- 4.40 In light of your answer to Problem 4.39, draw the two chair conformations of 1,1,3-trimethylcyclohexane and estimate the amount of strain energy in each. Which conformation is favored?
- 4.41 We'll see in Chapter 11 that alkyl halides undergo an elimination reaction to yield alkenes on treatment with strong base. For example, chlorocyclohexane gives cyclohexene on reaction with  $\text{NaNH}_2$ .



If axial chlorocyclohexanes are generally more reactive than their equatorial isomers, which do you think would react faster, *cis*-1-*tert*-butyl-2-chlorocyclohexane or *trans*-1-*tert*-butyl-2-chlorocyclohexane? Explain.

- 4.42 We saw in Problem 4.20 that *cis*-decalin is less stable than *trans*-decalin. Assume that the 1,3-diaxial interactions in *trans*-decalin are similar to those in axial methylcyclohexane (one  $\text{CH}_3\text{-H}$  interaction costs 0.9 kcal/mol), and calculate the magnitude of the energy difference between *cis*- and *trans*-decalin.
- 4.43 Using molecular models as well as structural drawings, explain why *trans*-decalin is rigid and cannot ring-flip, whereas *cis*-decalin can easily ring-flip.
- 4.44 How many geometric isomers of 1,2,3,4,5,6-hexachlorocyclohexane are there? Draw the structure of the most stable isomer.
- 4.45 Here's a tough one. Explain the observation that the all-*cis* isomer of 4-*tert*-butylcyclohexane-1,3-diol reacts readily with acetone and an acid catalyst to form an acetal, but that other stereoisomers do not react.



An acetal

In formulating your answer, draw the stable chair conformations of all four stereoisomers and of the product acetal. Use molecular models for help.

- 4.46 Increased substitution around a bond leads to increased strain. Take the four substituted butanes listed below, for example. For each compound, sight along the C2-C3 bond and draw Newman projections of the most stable and least stable conformations. Use the data in Table 4.1 to assign strain energy values to each

conformation. Which of the eight conformations is most stable, and which is least stable?

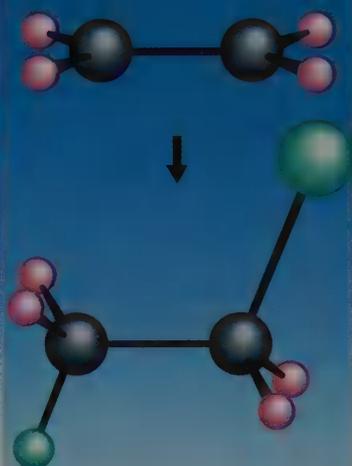
- (a) 2-Methylbutane  
(b) 2,2-Dimethylbutane  
(c) 2,3-Dimethylbutane  
(d) 2,2,3-Trimethylbutane

4.47 One of the two chair structures of *cis*-1-chloro-3-methylcyclohexane is more stable than the other by 3.7 kcal/mol. Which is it? What is the energy cost of a 1,3-diaxial interaction between a chlorine and a methyl group?

4.48 The German chemist J. Bredt proposed in 1935 that bicycloalkenes such as 1-norbornene that have a double bond to the bridgehead carbon are too strained to exist. Make a molecular model of 1-norbornene and explain Bredt's proposal.



1-Norbornene



## Chapter 5

# An Overview of Organic Reactions

When first approached, organic chemistry can seem like a bewildering collection of facts—a collection of millions of compounds, dozens of functional groups, and a vast number of reactions. With study, though, it becomes evident that there are only a few fundamental concepts that underlie *all* organic reactions.

Far from being a collection of isolated facts, organic chemistry is a beautifully logical subject unified by a few broad themes. When these themes are understood, learning organic chemistry becomes much easier and rote memorization can be avoided. The aim of this book is to point out the themes and to clarify the patterns that unify organic chemistry. We'll begin by taking an overview of the fundamental kinds of organic reactions that take place and seeing how they can be described.

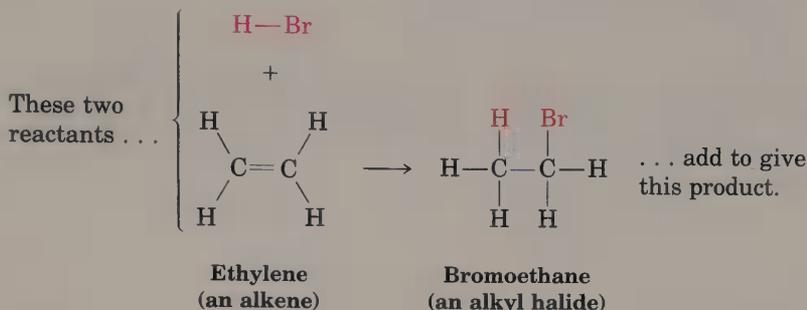
### 5.1 Kinds of Organic Reactions

Organic chemical reactions can be organized in two ways: by *what kinds* of reactions occur and by *how* reactions occur. When beginning a study of the subject, it's easier to look first at the kinds of reactions that take place. There are four particularly important kinds of organic reactions: additions, eliminations, substitutions, and rearrangements.

**Addition reactions** occur when two reactants add together to form a single new product with no atoms “left over.” We can generalize the process as:



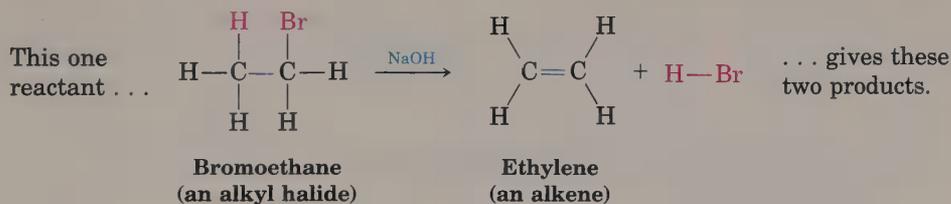
As an example of an important addition reaction that we'll be studying soon, alkenes such as ethylene react with acids such as HBr to yield alkyl halides:



**Elimination reactions** are, in a sense, the opposite of addition reactions. Eliminations occur when a single reactant splits apart into two products, a process we can generalize as:



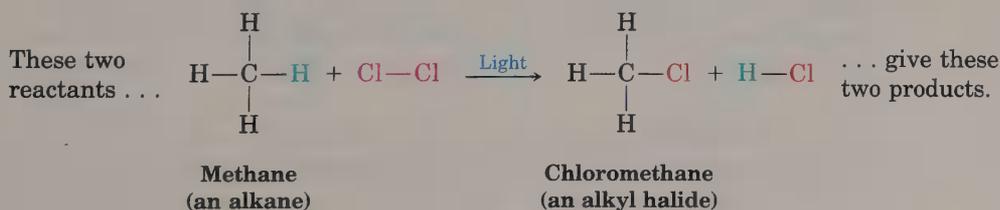
As an example of an important elimination reaction, alkyl halides such as bromoethane split apart into an acid and an alkene when treated with base:



**Substitution reactions** occur when two reactants exchange parts to give two new products, a process we can generalize as:



As an example of a substitution reaction, alkanes such as methane react with chlorine gas in the presence of ultraviolet light to yield alkyl chlorides. A -Cl group from chlorine substitutes for an -H group of methane, and two new products result:



**Rearrangement reactions** occur when a single reactant undergoes a reorganization of bonds and atoms to yield an isomeric product, a process we can generalize as:

This single reactant . . .  $A \longrightarrow B$  . . . gives this isomeric product.

As an example of a rearrangement reaction, the alkene 1-butene can be converted into its constitutional isomer, 2-butene, by treatment with an acid catalyst:



PROBLEM .....

5.1 Classify these reactions as additions, eliminations, substitutions, or rearrangements:

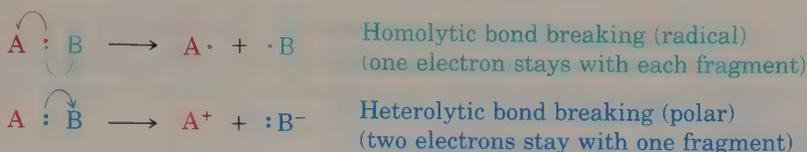
- (a)  $\text{CH}_3\text{Br} + \text{KOH} \longrightarrow \text{CH}_3\text{OH} + \text{KBr}$   
 (b)  $\text{CH}_3\text{CH}_2\text{OH} \longrightarrow \text{H}_2\text{C}=\text{CH}_2 + \text{H}_2\text{O}$   
 (c)  $\text{H}_2\text{C}=\text{CH}_2 + \text{H}_2 \longrightarrow \text{CH}_3\text{CH}_3$
- .....

## 5.2 How Organic Reactions Occur: Mechanisms

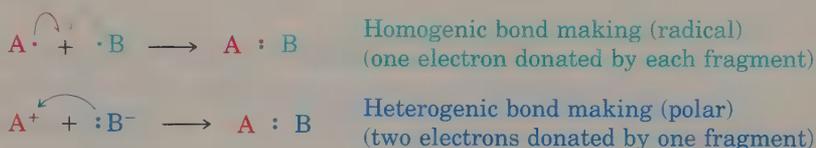
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Having looked at the kinds of reactions that take place, let's now see how reactions occur. An overall description of how a specific reaction occurs is called a reaction **mechanism**. A mechanism describes in detail exactly what takes place at each stage of a chemical transformation. It describes which bonds are broken and in what order, which bonds are formed and in what order, and what the relative rates of the steps are. A complete mechanism must also account for all reactants used, all products formed, and the amounts of each.

All chemical reactions involve bond breaking and bond making. When two reactants come together, react, and yield products, specific chemical bonds in the starting materials are broken, and specific chemical bonds in the products are formed. Fundamentally there are two ways that a covalent two-electron bond can break: A bond can break in an electronically *symmetrical* way so that one electron remains with each product fragment, or a bond can break in an electronically *unsymmetrical* way so that both bonding electrons remain with one product fragment, leaving the other fragment with an empty orbital. The symmetrical cleavage is called a **homolytic** process, and the unsymmetrical cleavage is called a **heterolytic** process.



Conversely, there are two ways that a covalent two-electron bond can form: A bond can form in an electronically symmetrical (**homogenic**) way when one electron is donated to the new bond by each reactant, or a bond can form in an electronically unsymmetrical (**heterogenic**) way when both bonding electrons are donated to the new bond by one reactant.



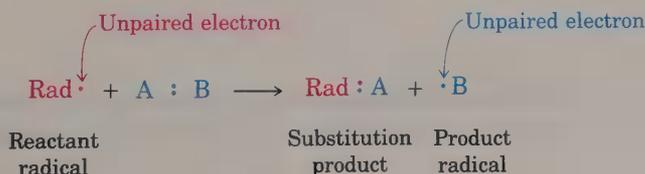
Those processes that involve symmetrical bond breaking and bond making are called **radical reactions**. A **radical** (sometimes called a “free radical”) is a species that contains an odd number of valence electrons and thus has a single, unpaired electron in one of its orbitals. Those processes that involve unsymmetrical bond breaking and bond making are called **polar reactions**. Polar reactions involve species that contain an even number of valence electrons and thus have only electron pairs in their orbitals. Polar processes are the more commonly encountered reaction type in organic chemistry, and a large part of this book is devoted to their description.

In addition to polar and radical reactions, there’s a third, less commonly encountered type called *pericyclic reactions*. Rather than explain pericyclic reactions now, though, we’ll study them in more detail at a later point.

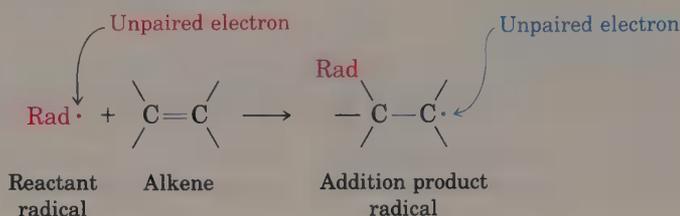
## 5.3 Radical Reactions and How They Occur

Radical reactions are not as common as polar reactions, but they are nevertheless quite important in organic chemistry, particularly in certain industrial processes. Let’s see how they occur.

Although most radicals are electrically neutral, they are highly reactive because they contain an atom with an odd number of electrons (usually seven) in its outer shell, rather than a stable noble-gas octet. They can achieve the desired octet in several ways. For example, a radical might abstract an atom from another molecule, leaving behind a new radical. The net result is a radical *substitution* reaction:



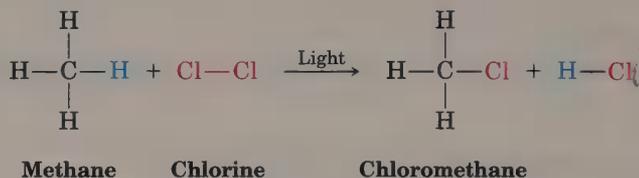
Alternatively, a reactant radical might add to an alkene, taking one electron from the alkene double bond and yielding a new radical. The net result is a radical *addition* reaction:



Let's look at a specific example of a radical reaction to see its characteristics. The details of the reaction aren't important at this point; it's important only to see that radical reactions involve odd-electron species.

## 5.4 An Example of a Radical Reaction: Chlorination of Methane

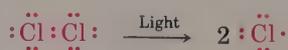
The chlorination of methane is a typical radical substitution reaction. Although inert to most reagents, alkanes react readily with chlorine to give chlorinated alkane products:



A more detailed discussion of this radical substitution reaction is given in Chapter 10. For the present, it's necessary only to know that studies have shown alkane chlorination to be a multistep process involving radicals.

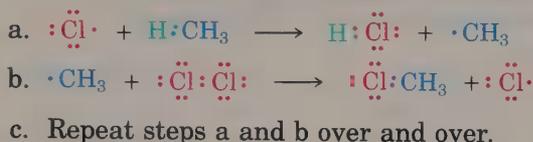
Radical substitution reactions normally require three kinds of steps: an initiation step, propagation steps, and termination steps.

1. *Initiation*—the initial production of radicals The initiation step begins the reaction by producing reactive radicals. In the present case, the relatively weak Cl–Cl bond is homolytically broken by irradiation with ultraviolet light. Two reactive chlorine radicals are produced, and further chemistry ensues.

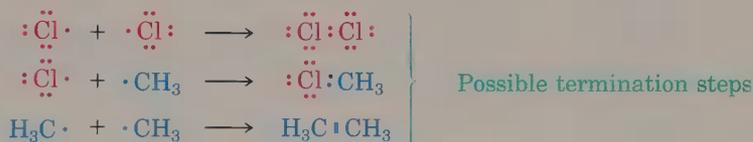


2. *Propagation steps*—radicals undergo substitution reactions Once chlorine radicals have been produced in small amounts, propagation

steps take place. When a reactive chlorine radical collides with a methane molecule, it abstracts a hydrogen atom to produce HCl and a methyl radical ( $\cdot\text{CH}_3$ ). This methyl radical reacts further with  $\text{Cl}_2$  in another substitution step to give the product chloromethane and a chlorine radical, which cycles back into the first propagation step. Once the sequence has been initiated, it becomes a self-sustaining cycle of repeating steps 2a and 2b, making the overall process a **chain reaction**.



3. *Termination steps—the chain is broken* Occasionally, two radicals might collide and combine to form a stable product in a termination step. When this happens, the reaction cycle is broken and the chain is ended. Such termination steps occur infrequently because the concentration of radicals in the reaction at any given moment is very small. Thus, the likelihood that two radicals will collide is also small.



Alkane chlorination is not a generally useful reaction because most alkanes (other than methane and ethane) have several different kinds of hydrogens, and mixtures of chlorinated products usually result in such cases. Nevertheless, radical chain reactions constitute a basic reaction type of considerable importance.

The radical substitution reaction just discussed is only one of several different processes that radicals can undergo. The fundamental principle behind all radical reactions is the same, however: All bonds are broken and formed by reaction of odd-electron species.

PROBLEM .....

- 5.2 When a mixture of methane and chlorine is irradiated, reaction commences immediately. When irradiation is stopped, the reaction gradually slows down but does not stop immediately. How do you account for this behavior?

PROBLEM .....

- 5.3 Draw and name all monochloro products you would expect to obtain from reaction of 2-methylpentane with chlorine.

## PROBLEM .....

- 5.4 Radical chlorination of pentane is a poor way to prepare 1-chloropentane,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{Cl}$ , but radical chlorination of neopentane,  $(\text{CH}_3)_4\text{C}$ , is a good way to prepare neopentyl chloride,  $(\text{CH}_3)_3\text{CCH}_2\text{Cl}$ . How do you account for this difference?
- .....

## 5.5 Polar Reactions and How They Occur

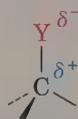
Polar reactions occur as the result of attractive forces between positive and negative charges on molecules. In order to see how these reactions take place, we need first to recall our previous discussion of polar covalent bonds and then to look more deeply into the effects of bond polarity on organic molecules.

Most organic molecules are electrically neutral; they have no net charge, either positive or negative. We saw in Section 2.4, however, that certain bonds within a molecule, particularly the bonds in functional groups, are polar. Bond polarity is a consequence of an unsymmetrical electron distribution in the bond and is due to the electronegativity of the atoms involved (Table 5.1, which repeats the information in Figure 2.3 for convenience).

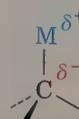
**TABLE 5.1** Relative Electronegativities of Some Common Elements

1A								
H	2A		3A	4A	5A	6A	7A	
2.1			B	C	N	O	F	
Li	Be		2.0	2.5	3.0	3.5	4.0	
1.0	1.5		Al	Si	P	S	Cl	
Na	Mg		1.5	1.8	2.1	2.5	3.0	
0.9	1.2						Br	
K							2.8	
0.8							I	
							2.5	

Atoms such as oxygen, nitrogen, fluorine, chlorine, and bromine are more electronegative than carbon. Thus, a carbon atom bonded to one of these electronegative atoms has a partial positive charge ( $\delta^+$ ), and the electronegative atom has a slight negative charge ( $\delta^-$ ). Conversely, metallic elements are less electronegative than carbon. A carbon atom bonded to a metal therefore has a partial *negative* charge, and the metal has a partial positive charge:



where  $Y = \text{O, N, Cl, Br, I}$



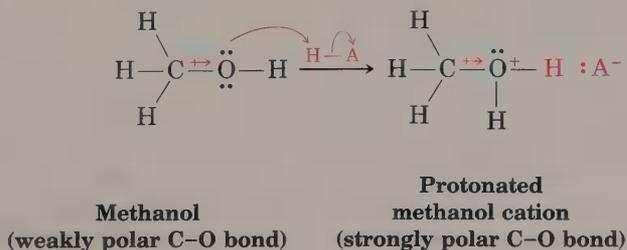
where  $M = \text{a metal such as Mg or Li}$

The polarity patterns of some common functional groups are shown in Table 5.2.

TABLE 5.2 Polarity Patterns in Some Common Functional Groups

Compound type	Functional group structure	Compound type	Functional group structure
Alcohol	$\begin{array}{c} \delta^+ \quad \delta^- \\   \quad   \\ -\text{C}-\text{OH} \\   \end{array}$	Carbonyl	$\begin{array}{c} \delta^+ \quad \delta^- \\   \quad   \\ -\text{C}=\text{O} \\   \end{array}$
Alkene	$\begin{array}{c} \diagup \quad \diagdown \\ \text{C}=\text{C} \\ \diagdown \quad \diagup \end{array}$ Symmetrical, nonpolar	Carboxylic acid	$\begin{array}{c} \text{O} \delta^- \\ // \\ -\text{C} \delta^+ \\   \\ \text{OH} \delta^- \end{array}$
Alkyl halide	$\begin{array}{c} \delta^+ \quad \delta^- \\   \quad   \\ -\text{C}-\text{X} \\   \end{array}$	Carboxylic acid chloride	$\begin{array}{c} \text{O} \delta^- \\ // \\ -\text{C} \delta^+ \\   \\ \text{Cl} \delta^- \end{array}$
Amine	$\begin{array}{c} \delta^+ \quad \delta^- \\   \quad   \\ -\text{C}-\text{NH}_2 \\   \end{array}$	Aldehyde	$\begin{array}{c} \text{O} \delta^- \\ // \\ -\text{C} \delta^+ \\   \\ \text{H} \end{array}$
Ether	$\begin{array}{c} \delta^+ \quad \delta^- \quad \delta^+ \\   \quad   \quad   \\ -\text{C}-\text{O}-\text{C}- \\   \quad   \end{array}$	Ester	$\begin{array}{c} \text{O} \delta^- \\ // \\ -\text{C} \delta^+ \\   \\ \text{O} \delta^- - \text{C} \end{array}$
Nitrile	$\begin{array}{c} \delta^+ \quad \delta^- \\   \quad   \\ -\text{C} \equiv \text{N} \\   \end{array}$	Ketone	$\begin{array}{c} \text{O} \delta^- \\ // \\ -\text{C} \delta^+ \\   \\ \text{C} \end{array}$
Grignard reagent	$\begin{array}{c} \delta^- \quad \delta^+ \\   \quad   \\ -\text{C}-\text{MgBr} \\   \end{array}$		
Alkyl lithium	$\begin{array}{c} \delta^- \quad \delta^+ \\   \quad   \\ -\text{C}-\text{Li} \\   \end{array}$		

This discussion of bond polarity is oversimplified in that we've considered only bonds that are *inherently* polar due to electronegativity effects. Polar bonds can also result from the interaction of functional groups with solvents and with acids. For example, the polarity of the carbon–oxygen bond in methanol is greatly enhanced by protonation:

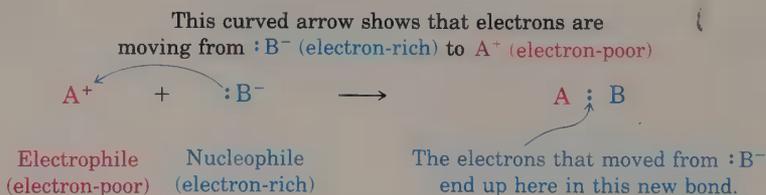


In neutral methanol, the carbon atom is somewhat electron-poor because the electronegative oxygen attracts carbon–oxygen bond electrons. In the protonated methanol cation, though, a full positive charge on oxygen *strongly* attracts electrons in the carbon–oxygen bond and makes the carbon much more electron-poor.

Yet a further consideration is the *polarizability* (as opposed to polarity) of an atom. As the electric field around a given atom changes due to changing interactions with solvent or with other polar reagents, the electron distribution around that atom also changes. The measure of this response to an external influence is the **polarizability** of the atom. Larger atoms with more loosely held electrons are more polarizable than smaller atoms with tightly held electrons. Thus, iodine is much more polarizable than fluorine. This means that the carbon–iodine bond, although electronically symmetrical according to the electronegativity table (Table 5.1), can nevertheless react as if it were polar.

What does functional-group polarity mean with respect to chemical reactivity? *Since unlike charges attract, the fundamental characteristic of all polar organic reactions is that electron-rich sites in the functional groups of one molecule react with electron-poor sites in the functional groups of another molecule.* Bonds are made when the electron-rich reagent donates a pair of electrons to the electron-poor reagent, and bonds are broken when one of the two product fragments leaves with the electron pair.

As we saw in Section 2.8, chemists normally indicate the electron-pair movement that occurs during a polar reaction by using curved arrows. A curved arrow indicates where electrons move as bonds in reactants are broken and bonds in products are formed. It means that an electron pair moves from the atom at the tail of the arrow to the atom at the head of the arrow during the reaction.



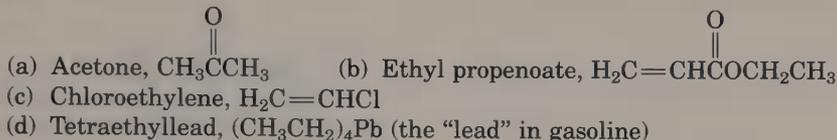
In referring to the fundamental polar process and to the species involved, chemists have coined the words *nucleophile* and *electrophile*. A **nucleophile** is a reagent that is “nucleus-loving”; nucleophiles have electron-rich sites and can form a bond by donating a pair of electrons to an electron-poor site. Nucleophiles are often, though not always, negatively charged. An **electrophile**, by contrast, is “electron-loving”; electrophiles have electron-poor sites and can form a bond by accepting a pair of electrons from a nucleophile. Electrophiles are often, though not always, positively charged.

If the definitions of nucleophiles and electrophiles sound similar to those given in Section 2.8 for Lewis acids and Lewis bases, that’s because there is indeed a correlation between electrophilicity–nucleophilicity and Lewis acidity–basicity. Lewis bases are electron donors and usually behave as nucleophiles, whereas Lewis acids are electron acceptors and usually behave as electrophiles. The major difference is that the terms *electrophile* and

*nucleophile* are normally used only when bonds to *carbon* are involved. We'll explore these ideas in more detail in Chapter 10.

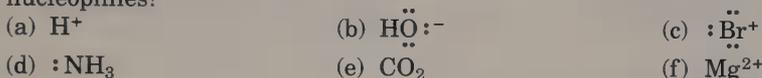
## PROBLEM.....

- 5.5 Identify the functional groups present in these molecules and show the direction of polarity in each.



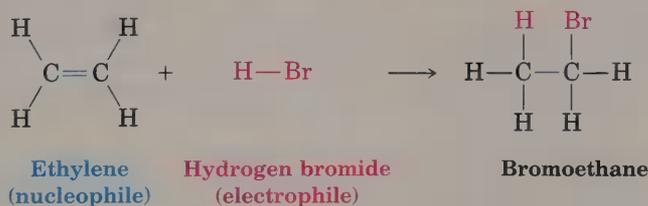
## PROBLEM.....

- 5.6 Which of the following would you expect to behave as electrophiles and which as nucleophiles?



## 5.6 An Example of a Polar Reaction: Addition of HBr to Ethylene

Let's look at a typical polar process, the addition reaction of ethylene with HBr that we saw earlier. When ethylene is treated with hydrogen bromide at room temperature, bromoethane is produced. Overall, the reaction can be formulated as follows:



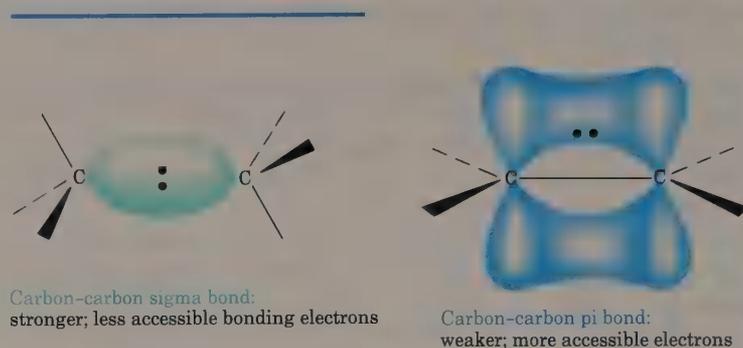
This reaction, an example of a general polar reaction type known as an *electrophilic addition*, can be understood in terms of the general concepts just discussed. We'll begin by looking at the nature of the two reactants.

What do we know about ethylene? We know from Section 1.10 that a carbon-carbon double bond results from orbital overlap of two  $sp^2$ -hybridized carbon atoms. The sigma part of the double bond results from  $sp^2$ - $sp^2$  overlap, and the pi part results from  $p$ - $p$  overlap.

What kind of chemical reactivity might we expect of carbon-carbon double bonds? We know that *alkanes*, such as ethane, are rather inert, because all valence electrons are tied up in strong, relatively nonpolar, carbon-carbon and carbon-hydrogen bonds. Furthermore, we know that the bonding electrons in alkanes are relatively inaccessible to external reagents because they are sheltered in sigma orbitals between nuclei.

The electronic situation in ethylene and other alkenes is quite different. For one thing, double bonds have greater electron density than single bonds

have—four electrons in a double bond versus only two electrons in a single bond. Equally important, the electrons in the pi bond are accessible to external reagents because they are located above and below the plane of the double bond rather than being sheltered between the nuclei (Figure 5.1).



**Figure 5.1** A comparison of carbon–carbon single and double bonds. A double bond is both more electron-rich (nucleophilic) and more accessible to attack by external reagents than a single bond.

Both electron richness and electron accessibility lead to a prediction of high reactivity for carbon–carbon double bonds. In the terminology of polar reactions just introduced, we might predict that carbon–carbon double bonds should behave as *nucleophiles*. That is, the chemistry of alkenes should be dominated by reaction of the electron-rich double bond with electron-poor reagents. This is exactly what we find: The most important reaction of alkenes is their reaction with electrophiles.

What about HBr? As a strong mineral acid, HBr is a powerful proton ( $\text{H}^+$ ) donor. Since a proton is positively charged and electron-poor, it is a good electrophile. Thus, the reaction between  $\text{H}^+$  and ethylene is a typical electrophile–nucleophile combination, characteristic of all polar reactions. (Although chemists often talk about “ $\text{H}^+$ ” when referring to acids, in fact there’s really no such species. Protons are always associated with another molecule for stability, for example with water in  $\text{H}_3\text{O}^+$ .)

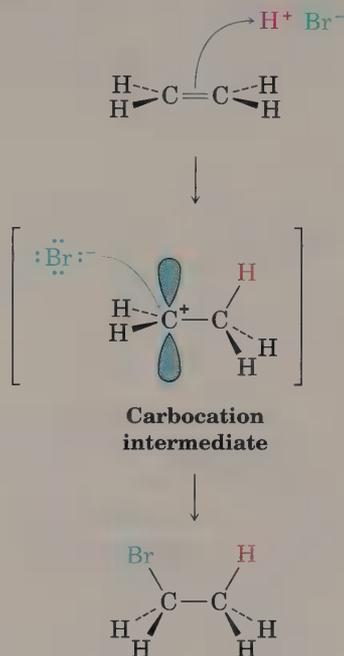
We’ll see more details about alkene electrophilic addition reactions shortly, but for the present we can look at the reaction as taking place by the pathway shown in Figure 5.2. The reaction begins when the alkene (ethylene) donates a pair of electrons from its  $\text{C}=\text{C}$  double bond to form a new single bond to  $\text{H}^+$ . Note that the curved arrow in the first step of Figure 5.2 indicates an electron pair moving from the double bond to form a single bond to hydrogen.

One of the double-bond carbon atoms forms a new bond to the incoming hydrogen, but the other carbon atom, having lost its share of the double-bond electrons, is now trivalent, has only six valence electrons, and is left with the positive charge. This positively charged species, a carbon cation or **carbocation**, is itself an electrophile that can accept an electron pair from

nucleophilic bromide anion to form a C–Br bond and give neutral bromoethane addition product. Once again, a curved arrow in Figure 5.2 shows the electron-pair movement from bromide anion to carbon.

The electrophile  $\text{H}^+$  is attacked by the pi electrons of the double bond, and a new C–H sigma bond is formed. This leaves the other carbon atom with a + charge and a vacant  $p$  orbital.

$\text{Br}^-$  donates an electron pair to the positively charged carbon atom, forming a C–Br sigma bond and yielding the neutral addition product.

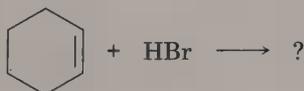


**Figure 5.2** The polar addition reaction of HBr and ethylene. The reaction takes place in two steps, both of which involve electrophile–nucleophile interactions.

The electrophilic addition of HBr to ethylene is only one example of a polar process; there are many other types that we'll study in detail in later chapters. Regardless of the details of individual reactions, all polar processes can be accounted for in the general terms just presented. To repeat: Polar reactions take place between an electron-poor site and an electron-rich site and involve the sharing of an electron pair from a nucleophile with an electrophile.

**PROBLEM** .....

**5.7** What product would you expect from reaction of HBr with cyclohexene?



.....

## 5.7 Describing a Reaction: Rates and Equilibria

All chemical reactions can be written as equilibrium processes. Starting materials can react to give products, and products can revert back to starting materials. The position of a chemical equilibrium is expressed by an equation in which  $K_{\text{eq}}$ , the equilibrium constant, is equal to the product concentrations multiplied together, divided by the starting material concentrations multiplied together. For the generalized reaction

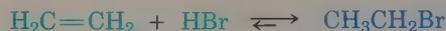


we have

$$K_{\text{eq}} = \frac{[\text{Products}]}{[\text{Reactants}]} = \frac{[C]^c[D]^d}{[A]^a[B]^b}$$

The equilibrium constant tells which side of the reaction arrow is energetically more favored. If  $K_{\text{eq}}$  is larger than 1, then the product concentrations  $[C]^c[D]^d$  are larger than the reactant concentrations  $[A]^a[B]^b$ , and the reaction proceeds as written from left to right. Conversely, if  $K_{\text{eq}}$  is smaller than 1, the reaction does not take place as written.

In the reaction of ethylene with HBr, for example, we can write the following equilibrium expression and can determine experimentally that the equilibrium constant at room temperature is approximately  $7.5 \times 10^7$ :

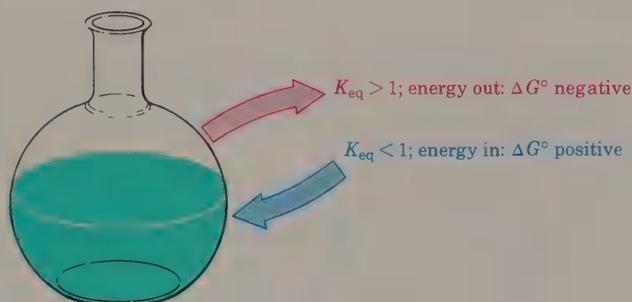


$$K_{\text{eq}} = \frac{[\text{CH}_3\text{CH}_2\text{Br}]}{[\text{HBr}][\text{H}_2\text{C}=\text{CH}_2]} = 7.5 \times 10^7$$

Since  $K_{\text{eq}}$  is relatively large, the reaction proceeds in the direction written and 99.999 997% of all ethylene molecules are converted into bromoethane molecules. Although we often speak of such reactions as “going to completion,” this is imprecise terminology since in practically no reaction does every molecule react. For practical purposes, though, equilibrium constants of greater than  $10^3$  can be considered to indicate “complete” reaction, since the amount of reactant left will be barely detectable (less than 0.1%).

What determines the direction in which a reaction proceeds? *In order for a reaction to have a favorable equilibrium constant and to proceed as written, the energy level of the products must be lower than the energy level of the starting materials.* In other words, energy must be given off. The situation is analogous to that of a rock poised near the top of a hill. The rock is in a high-energy position because energy was required to raise it to the top. When it rolls downhill, it releases energy until it reaches a more stable low-energy position at the bottom.

The total amount of energy change during a reaction is called the **standard Gibbs free-energy change**,  $\Delta G^\circ$ .<sup>1</sup> (The Greek letter delta,  $\Delta$ , is the mathematical symbol for the difference between two numbers, in this case the difference between the free energy of the products and the free energy of the starting materials.) For a favorable reaction,  $\Delta G^\circ$  has a negative value, meaning that energy is released *to* the surroundings; for an unfavorable reaction,  $\Delta G^\circ$  has a positive value, meaning that energy is absorbed *from* the surroundings.



Because the equilibrium constant,  $K_{\text{eq}}$ , and the free-energy change,  $\Delta G^\circ$ , both measure whether or not a reaction is favored, they are mathematically related:

$$\Delta G^\circ = \text{Free energy of products} - \text{Free energy of reactants}$$

$$\Delta G^\circ = -RT \ln K_{\text{eq}} \quad \text{or} \quad K_{\text{eq}} = e^{-\Delta G^\circ/RT}$$

where  $R = 1.986 \text{ cal/degree} \cdot \text{mol}$  ( $8.309 \text{ J/degree} \cdot \text{mol}$ ), the gas constant

$T = \text{Absolute temperature in kelvins}$

$e = 2.718$ , the base of natural logarithms

$\ln K_{\text{eq}} = \text{Natural logarithm of } K_{\text{eq}}$

As an example of how this mathematical relationship can be used, the reaction of ethylene with HBr has  $K_{\text{eq}} = 7.5 \times 10^7$ , and we can therefore calculate that  $\Delta G^\circ = -10.7 \text{ kcal/mol}$  ( $-45 \text{ kJ/mol}$ ) at room temperature (298 K):

$$K_{\text{eq}} = 7.5 \times 10^7 \quad \text{and} \quad \ln K_{\text{eq}} = 18.1$$

$$\Delta G^\circ = -RT \ln K_{\text{eq}} = -(1.986 \text{ cal/K} \cdot \text{mol})(298 \text{ K})(18.1)$$

$$= -10,700 \text{ cal/mol} = -10.7 \text{ kcal/mol}$$

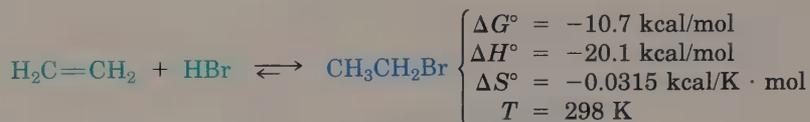
<sup>1</sup>The superscript  $^\circ$  in  $\Delta G^\circ$  means that all components are in their standard states. For gases, the standard state is at 1 atm pressure; for solutions, the standard state is 1 M in concentration. The superscript is dropped and the free energy change is  $\Delta G$  if a reaction is carried out under other than standard-state conditions.

To what is the energy change during a reaction due? The Gibbs free-energy change is attributable to a combination of two factors, an **enthalpy** factor,  $\Delta H^\circ$ , and an **entropy** factor,  $\Delta S^\circ$ :

$$\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$$

where  $T$  is the absolute temperature in kelvins

For the reaction of ethylene with HBr at room temperature (298 K), the values are  $\Delta G^\circ = -10.7$  kcal/mol,  $\Delta H^\circ = -20.1$  kcal/mol, and  $\Delta S^\circ = -31.5$  cal/K · mol:



The enthalpy term,  $\Delta H^\circ$ , is called the **heat of reaction** and is a measure of the change in total bonding energy during a reaction. If  $\Delta H^\circ$  is negative, as in the reaction of HBr with ethylene, then the bonds in the product are stronger (more stable) than the bonds in the starting material, heat is evolved, and the reaction is said to be **exothermic**. If  $\Delta H^\circ$  is positive, the bonds in the product are weaker (less stable) than the bonds in the starting material, heat is absorbed, and the reaction is said to be **endothermic**. For example, if a certain reaction breaks reactant bonds with a total strength of 100 kcal/mol and forms new product bonds with a total strength of 120 kcal/mol, then  $\Delta H^\circ$  for the reaction is  $-20$  kcal/mol and the reaction is exothermic. [*Remember*: Breaking bonds takes energy, and making bonds releases energy.]

$$\text{Strength of bonds broken: } \Delta H^\circ = 100 \text{ kcal/mol}$$

$$\text{Strength of bonds formed: } \Delta H^\circ = -120 \text{ kcal/mol}$$

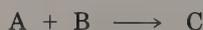
$$\text{Net change} = -20 \text{ kcal/mol}$$

The entropy term,  $\Delta S^\circ$ , is a measure of the change in the amount of molecular disorder or freedom of motion caused by a reaction. To illustrate with an elimination reaction of the type



there is more freedom of movement (disorder) in the products than in the reactant because one molecule has split into two. Thus, there is a net gain in entropy during the reaction, and  $\Delta S^\circ$  has a positive value.

On the other hand, for an addition reaction of the type



the opposite is true. Because such reactions restrict the freedom of movement of two molecules by joining them together, the product has *less* disorder

than the reactants, and  $\Delta S^\circ$  has a negative value. The reaction of ethylene and HBr to yield bromoethane is such a process ( $\Delta S^\circ = -31.5 \text{ cal/K} \cdot \text{mol}$ ).

Of the two terms that make up  $\Delta G^\circ$ , the enthalpy term ( $\Delta H^\circ$ ) is usually larger and more important than the entropy term ( $T\Delta S^\circ$ ) at normal reaction temperatures. Furthermore, enthalpy changes during reactions are relatively easily measured, and large compilations of data are available. For these reasons, the entropy contribution,  $T\Delta S^\circ$ , is sometimes ignored when making thermodynamic arguments, and chemists often make the simplifying assumption that  $\Delta G^\circ \approx \Delta H^\circ$ . Table 5.3 describes the thermodynamic terms more fully.

**TABLE 5.3** Explanation of Thermodynamic Quantities:  $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$

Term	Name	Explanation
$\Delta G^\circ$	Gibbs free-energy change (kcal/mol)	Overall energy difference between reactants and products. When $\Delta G^\circ$ is negative, a reaction can occur spontaneously. $\Delta G^\circ$ is related to the equilibrium constant by the equation $\Delta G^\circ = -RT \ln K_{\text{eq}}$
$\Delta H^\circ$	Enthalpy change (kcal/mol)	Heat of reaction; the energy difference between strengths of bonds broken in a reaction and bonds formed
$\Delta S^\circ$	Entropy change (cal/degree $\cdot$ mol)	Overall change in freedom of motion or "disorder" resulting from reaction; usually much smaller than $\Delta H^\circ$

Knowledge of equilibria and of the simple thermodynamic quantities that go into making up an equilibrium are extremely valuable, but it's important to realize their limitations. What an equilibrium constant tells is the *position* of the equilibrium, or how much product is formed. What the equilibrium constant does *not* tell is the *rate* of reaction, or how fast the equilibrium is established. Some reactions are extremely slow even though they have favorable equilibrium constants. Gasoline is stable when stored, for example, because the rate of its reaction with oxygen is slow under normal circumstances. Under other circumstances, however, such as contact with a lighted match, gasoline reacts rapidly with oxygen and undergoes complete conversion to the equilibrium products water and carbon dioxide. Rates (*how fast* a reaction occurs) and equilibria (*how much* a reaction occurs) are entirely different.

**Rate**  $\longrightarrow$  Is the reaction fast or slow?

**Equilibrium**  $\longrightarrow$  In what direction does the reaction proceed?

**PROBLEM** .....

- 5.8 Which reaction is more favored, one with  $\Delta G^\circ = -11 \text{ kcal/mol}$  or one with  $\Delta G^\circ = +11 \text{ kcal/mol}$ ?

PROBLEM .....

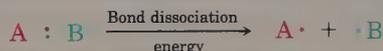
- 5.9 Which reaction is more exothermic, one with
- $K_{\text{eq}} = 1000$
- or one with
- $K_{\text{eq}} = 0.001$
- ?

PROBLEM .....

- 5.10 Use your calculator to find the value of
- $\Delta G^\circ$
- at 298 K for reactions where
- $K_{\text{eq}} = 1000$
- ,
- $K_{\text{eq}} = 1$
- , and
- $K_{\text{eq}} = 0.001$
- . Then find the value of
- $K_{\text{eq}}$
- for reactions where
- $\Delta G^\circ = -10$
- kcal/mol,
- $\Delta G^\circ = 0$
- kcal/mol, and
- $\Delta G^\circ = +10$
- kcal/mol.

## 5.8 Describing a Reaction: Bond Dissociation Energies

We've just seen that energy is released (negative  $\Delta H^\circ$ ) when a bond is made and absorbed (positive  $\Delta H^\circ$ ) when a bond is broken. The measure of the energy change on bond making or bond breaking is a quantity called the **bond dissociation energy**. Bond dissociation energy is defined as the amount of energy required to break a given bond to produce two radical fragments when the molecule is in the gas phase at 25°C.

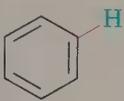
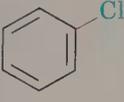
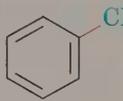
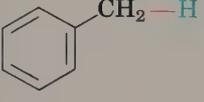
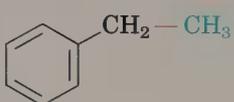
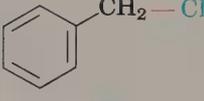
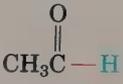
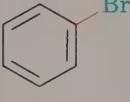
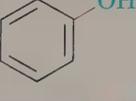
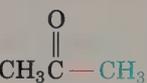


Each specific bond has its own characteristic strength, and extensive tabulations of bond-strength data are available. For example, a C–H bond in methane has a bond dissociation energy  $\Delta H^\circ = +104$  kcal/mol (435 kJ/mol), meaning that 104 kcal/mol is required to break a C–H bond of methane to give the two radical fragments  $\cdot\text{CH}_3$  and  $\cdot\text{H}$ . Conversely, 104 kcal/mol of energy is *released* when a methyl radical and a hydrogen atom combine to form methane. Table 5.4 lists some other bond-strength data.

If enough bond dissociation energies were known, it would seem possible to calculate  $\Delta H^\circ$  for any hypothetical process, thereby avoiding a lot of time-consuming work in the laboratory. Unfortunately, there are two problems with using bond-strength data to calculate whether or not a reaction is favorable. First, the calculation says nothing about the rate of reaction, since a reaction may have a favorable  $\Delta H^\circ$  and still not take place. Second, bond dissociation energies refer to reactions occurring in the gas phase and aren't directly relevant to solution chemistry.

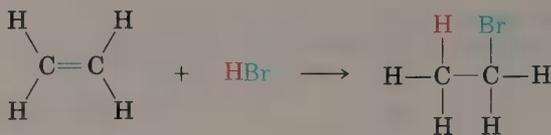
In practice, most organic reactions are carried out in solution, where solvent molecules can surround and interact with dissolved reagents, a phenomenon called *solvation*. Solvation can weaken bonds and cause large changes in the value of  $\Delta H^\circ$  for a given reaction. In addition, the entropy term,  $\Delta S^\circ$ , can also be affected by solvent molecules, since the solvation of polar reagents by polar solvents causes a certain amount of orientation (reduces the amount of disorder) in the solvent. Although we can often use

TABLE 5.4 Bond Dissociation Energy Data for the Reaction  $A-B \rightarrow A \cdot + B \cdot$ 

Bond	$\Delta H^\circ$ (kcal/mol)	Bond	$\Delta H^\circ$ (kcal/mol)	Bond	$\Delta H^\circ$ (kcal/mol)
H—H	104	$(CH_3)_3C-I$	50	$CH_3-CH_3$	88
H—F	136	$H_2C=CH-H$	108	$C_2H_5-CH_3$	85
H—Cl	103	$H_2C=CH-Cl$	88	$(CH_3)_2CH-CH_3$	84
H—Br	88	$H_2C=CHCH_2-H$	87	$(CH_3)_3C-CH_3$	81
H—I	71	$H_2C=CHCH_2-Cl$	69	$H_2C=CH-CH_3$	97
Cl—Cl	58		112	$H_2C=CHCH_2-CH_3$	74
Br—Br	46		97		102
I—I	36		85		72
$CH_3-H$	104		70		86
$CH_3-Cl$	84		82	HO—H	119
$CH_3-Br$	70		112	HO—OH	51
$CH_3-I$	56	$HC\equiv C-H$	125	$CH_3O-H$	102
$CH_3-OH$	91			$CH_3S-H$	88
$CH_3-NH_2$	80			$C_2H_5O-H$	103
$C_2H_5-H$	98				77
$C_2H_5-Cl$	81			$CH_3CH_2O-CH_3$	81
$C_2H_5-Br$	68			$NH_2-H$	103
$C_2H_5-I$	53			H—CN	130
$C_2H_5-OH$	91				
$(CH_3)_2CH-H$	95				
$(CH_3)_2CH-Cl$	80				
$(CH_3)_2CH-Br$	68				
$(CH_3)_3C-H$	91				
$(CH_3)_3C-Cl$	79				
$(CH_3)_3C-Br$	65				

bond-strength data to get a rough idea of how thermodynamically favorable a given reaction might be, we have to keep in mind that the answer is only approximate.

To take a familiar example, what do the thermodynamics look like for the reaction of gaseous HBr with gaseous ethylene? By totaling the energy released in making new bonds and the energy absorbed in breaking old bonds, we can calculate an approximate  $\Delta H^\circ$  for the overall reaction.



<i>Bonds broken</i>		<i>Bonds formed</i>	
H—Br	$\Delta H^\circ = 87$ kcal/mol	C—H	$\Delta H^\circ = 99$ kcal/mol
$\frac{1}{2}$ C=C	$\Delta H^\circ = 62$ kcal/mol	C—Br	$\Delta H^\circ = 68$ kcal/mol
$\Delta H^\circ = 149$ kcal/mol		$\Delta H^\circ = 167$ kcal/mol	

$$\text{Net change:} = 149 - 167 = -18 \text{ kcal/mol}$$

Since the bonds formed in the gas-phase reaction of HBr with ethylene are stronger than the bonds broken, we calculate that the reaction is favorable by about  $-18$  kcal/mol ( $-75$  kJ/mol), in good agreement with the actual value of  $-20.1$  kcal/mol. We can't be *certain* that the reaction will actually take place, since the calculation says nothing about reaction rates, but it does, in fact, occur as written.

#### PROBLEM

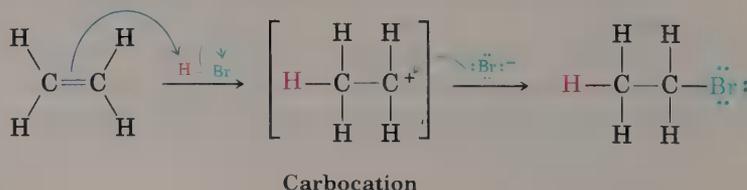
- 5.11** Calculate  $\Delta H^\circ$  for each step in the gas-phase radical substitution reaction of chlorine with methane (Section 5.4):
- $\text{Cl}_2 \longrightarrow 2 \text{Cl}\cdot$
  - $\text{CH}_4 + \text{Cl}\cdot \longrightarrow \cdot\text{CH}_3 + \text{HCl}$
  - $\cdot\text{CH}_3 + \text{Cl}_2 \longrightarrow \text{CH}_3\text{Cl} + \text{Cl}\cdot$
  - What is the overall  $\Delta H^\circ$  for the reaction? Consider only the propagation steps (b) and (c) in deciding your answer.

#### PROBLEM

- 5.12** Calculate  $\Delta H^\circ$  for these reactions:
- $\cdot\text{CH}_3\text{CH}_2\text{OCH}_3 + \text{HI} \longrightarrow \text{CH}_3\text{CH}_2\text{OH} + \text{CH}_3\text{I}$
  - $\text{CH}_3\text{Cl} + \text{NH}_3 \longrightarrow \text{CH}_3\text{NH}_2 + \text{HCl}$

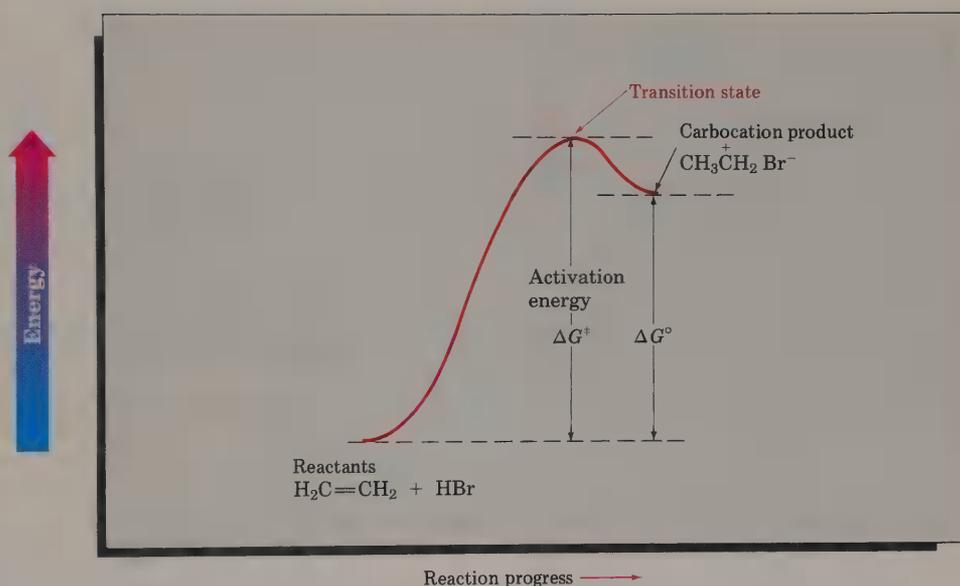
## 5.9 Describing a Reaction: Energy Diagrams and Transition States

In order for a reaction to take place, reactant molecules must collide, and reorganization of atoms and bonds must occur. Let's again look at the addition reaction of HBr with ethylene:



As the reaction proceeds, ethylene and HBr must approach each other, the ethylene pi bond and H–Br bond must break, a new carbon–hydrogen bond must form in the first step, and a new carbon–bromine bond must form in the second step.

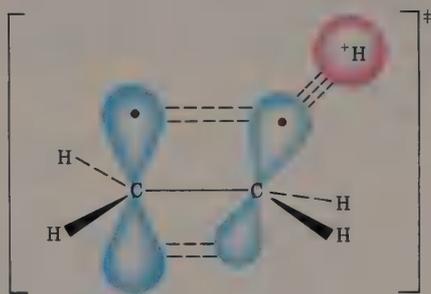
Over the years, chemists have developed a method for graphically depicting the energy changes that occur during a reaction using **reaction energy diagrams** of the sort shown in Figure 5.3. The vertical axis of the diagram represents the total energy of all reactants, while the horizontal axis, called the *reaction coordinate*, represents the progress of the reaction from beginning (left) to end (right). Let's take a careful look at the reaction, one step at a time, and see how the addition of HBr to ethylene can be described on a reaction energy diagram.



**Figure 5.3** A reaction energy diagram for the first step in the reaction of ethylene with HBr. The energy difference between reactants and transition state,  $\Delta G^\ddagger$ , controls the reaction rate. The energy difference between reactants and carbocation product,  $\Delta G^\circ$ , controls the position of the equilibrium.

At the beginning of the reaction, ethylene and HBr have the total amount of energy indicated by the reactant level on the left side of the diagram. As the two molecules crowd together and reaction commences, their electron clouds repel each other causing the energy level to rise. If the collision has occurred with sufficient force and proper orientation, the reactants continue to approach each other despite the rising repulsion until the new carbon–hydrogen bond starts to form. At some point, a structure of maximum energy is reached, a structure we call the **transition state**.

The transition state represents the highest-energy structure involved in the step, is unstable, and can't be isolated. Nevertheless, we can imagine it to be a kind of activated complex of the two reactants in which the carbon-carbon pi bond is partially broken and the new carbon-hydrogen bond is partially formed (Figure 5.4).



**Figure 5.4** A hypothetical transition-state structure for the first step of the reaction of ethylene with HBr. The C-C pi bond is just beginning to break, and the C-H bond is just beginning to form.

The energy difference between reactants and transition state is called the **activation energy**,  $\Delta G^\ddagger$ , and determines how rapidly the reaction occurs. (A double dagger superscript,  $\ddagger$ , always refers to the transition state.) A large activation energy corresponds to a large energy difference between reactants and transition state, and results in a slow reaction because few reacting molecules collide with enough energy to climb the high barrier. A small activation energy results in a rapid reaction because almost all reacting molecules are energetic enough to climb to the transition state.

The situation whereby reactants need enough energy to climb the activation barrier from starting material to transition state is like the situation of hikers who need enough energy to climb over a mountain pass. If the pass is a high one, the hikers need a lot of energy and will surmount the barrier very slowly. If the pass is low, however, the hikers need less energy and will reach the top quickly.

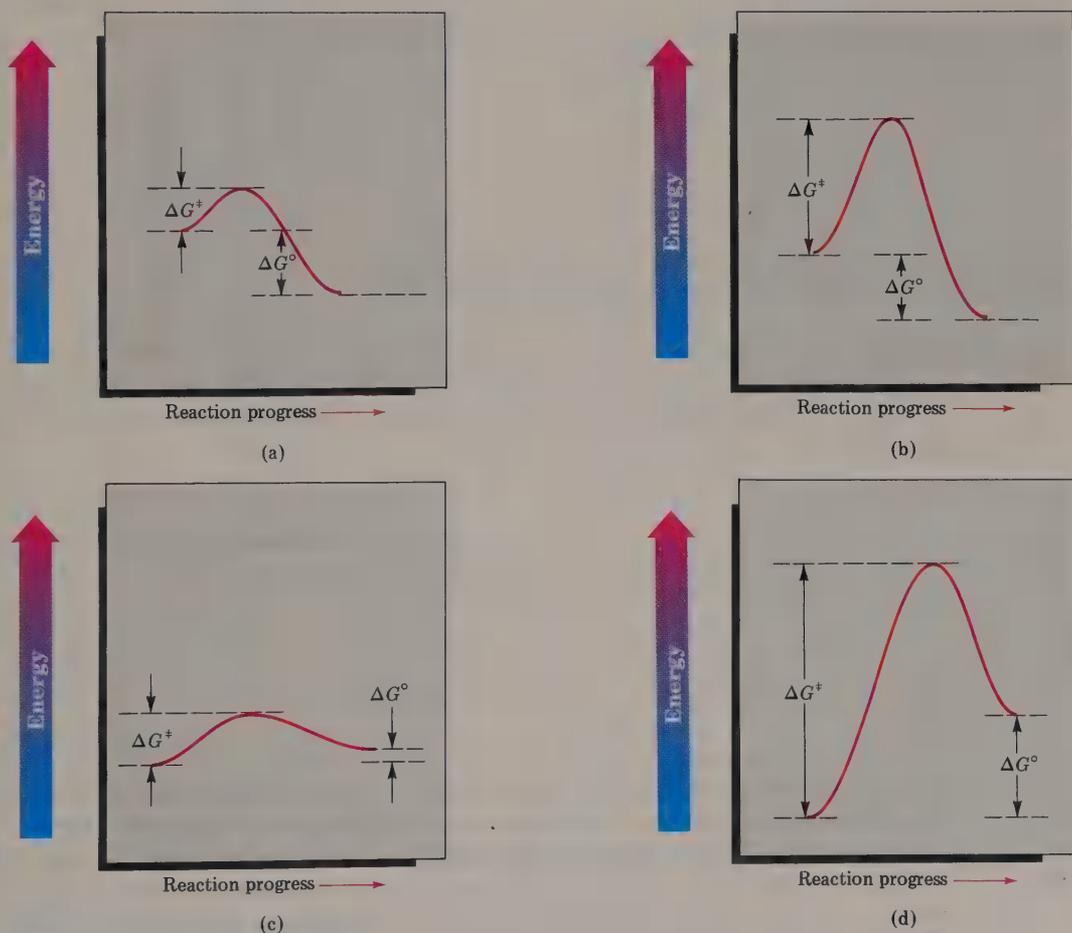
Although it's difficult to generalize, most organic reactions have activation energies in the range 10–35 kcal/mol (40–150 kJ/mol). The reaction of ethylene with HBr, for example, has an activation energy of approximately 34 kcal/mol. Reactions with activation energies less than 20 kcal/mol take place spontaneously at room temperature or below, whereas reactions with higher activation energies normally require heating. Heat provides the energy necessary for the reactants to climb the activation barrier.

Once the high-energy transition state has been reached, the reaction can either continue on to give the product or revert back to starting materials. Since both choices are energetically downhill from the high point of the transition state, both are equally likely. If reversion to starting materials occurs, no net change in the system is observed. If, however, the reaction

continues on to give carbocation, energy is released as the new C–H bond forms fully, and the curve on the reaction energy diagram therefore turns downward until it reaches a minimum.

This minimum point represents the energy level of the carbocation product of the first step. The energy change,  $\Delta G^\circ$ , between starting materials and carbocation is simply the difference between the two levels on the diagram. Since the carbocation is less stable than the starting alkene, the first step is endothermic, and energy is absorbed.

Not all reaction energy diagrams are like that for the reaction of ethylene and HBr. Each specific reaction has its own specific energy profile. Some reactions are very fast (low  $\Delta G^\ddagger$ ) and some are very slow (high  $\Delta G^\ddagger$ ); some have a negative  $\Delta G^\circ$  and some have a positive  $\Delta G^\circ$ . Figure 5.5 illustrates some different possibilities for energy profiles. Note in this figure the use



**Figure 5.5** Hypothetical reaction energy diagrams:

- (a) a fast exothermic reaction (small  $\Delta G^\ddagger$ , negative  $\Delta G^\circ$ );
- (b) a slow exothermic reaction (large  $\Delta G^\ddagger$ , negative  $\Delta G^\circ$ );
- (c) a fast endothermic reaction (small  $\Delta G^\ddagger$ , small positive  $\Delta G^\circ$ );
- (d) a slow endothermic reaction (large  $\Delta G^\ddagger$ , positive  $\Delta G^\circ$ ).

of the words *exothermic* and *endothermic* to refer to reactions in which  $\Delta G^\circ$  is negative and positive, respectively. This usage is not strictly correct, since these words refer only to  $\Delta H^\circ$  and not to  $\Delta G^\circ$ . As mentioned earlier, though, chemists often make the simplifying assumption that  $\Delta G^\circ$  and  $\Delta H^\circ$  are approximately equal.

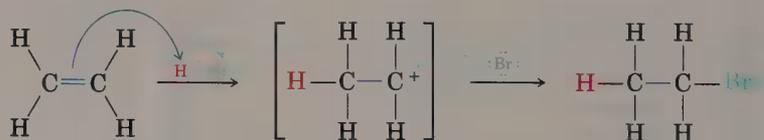
It should be emphasized again that the overall energy change occurring during a reaction is measured by  $\Delta G^\circ$ , not  $\Delta H^\circ$ . Calculations of changes in bond dissociation energies ( $\Delta H^\circ$ ) such as that done earlier for the reaction of ethylene and HBr (Section 5.8) are useful but can give only an indication as to whether or not a given reaction will have a favorable equilibrium constant. Such calculations don't take solvent effects or entropy factors into account.

PROBLEM.....

- 5.13 Which reaction is faster: one with  $\Delta G^\ddagger = 45$  kJ/mol or one with  $\Delta G^\ddagger = 70$  kJ/mol? Is it possible to predict which of the two has the larger  $K_{\text{eq}}$ ?
- .....

## 5.10 Describing a Reaction: Intermediates

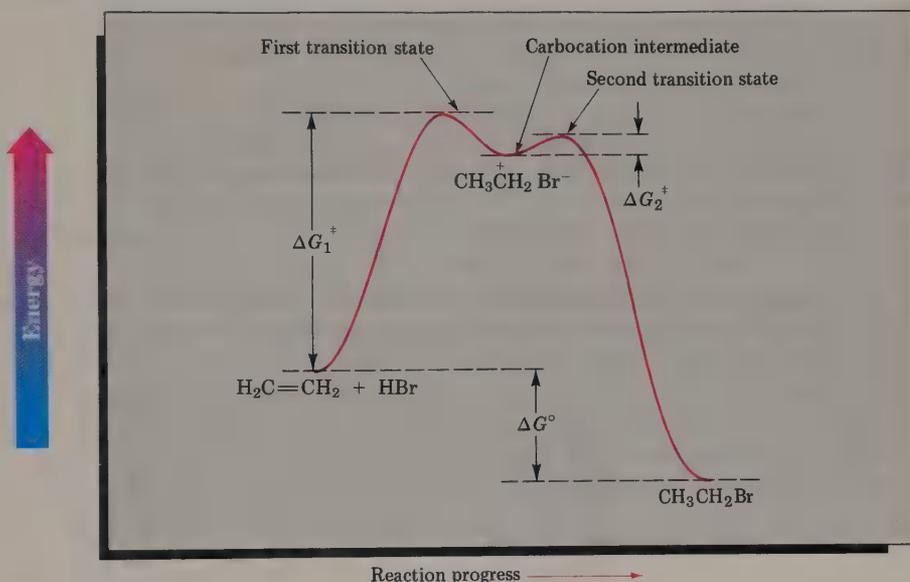
How can we characterize the carbocation structure formed in the first step of the reaction of ethylene with HBr? The carbocation is clearly different from the starting materials, yet it isn't a transition state and it isn't a final product.



Reaction intermediate

We call the carbocation, which is formed briefly during the course of the multistep reaction, a **reaction intermediate**. As soon as the intermediate is formed in the first step by reaction of ethylene with  $\text{H}^+$ , it reacts further with bromide ion in a second step to give the final product, bromoethane. This second step has its own activation energy,  $\Delta G^\ddagger$ , its own transition state, and its own energy change,  $\Delta G^\circ$ . We can picture the second transition state as an activated complex between the electrophilic carbocation intermediate and the nucleophilic bromide anion, in which the bromide ion is donating a pair of electrons to the positively charged carbon atom and the new C-Br bond is just starting to form.

A complete energy diagram for the overall reaction of ethylene with HBr is shown in Figure 5.6. In essence, we draw diagrams for each of the individual steps and then join them in the middle so that the *product* of step 1 (the carbocation) serves as the *starting material* for step 2. As indicated in Figure 5.6, the reaction intermediate lies at an energy minimum between steps 1 and 2. Since the energy level of this intermediate is higher than the level of either starting material (ethylene + HBr) or product (bromoethane), the intermediate is reactive and cannot be isolated. It is, however, more stable than either of the two transition states that surround it.

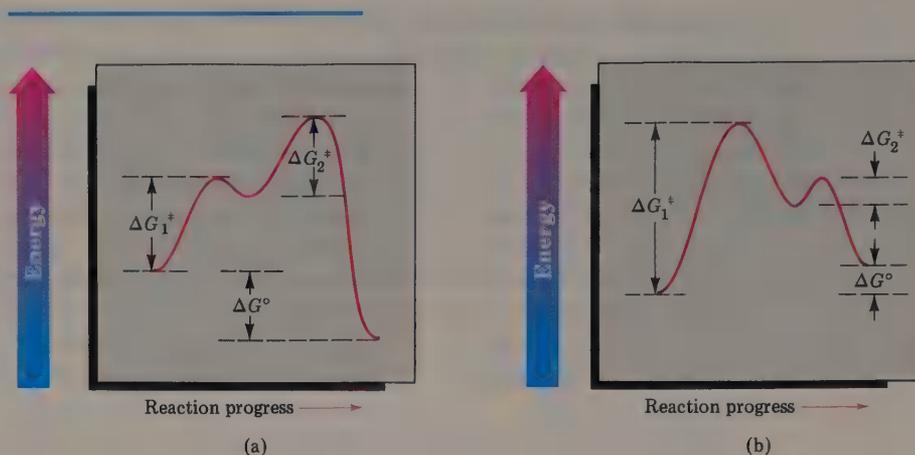


**Figure 5.6** An overall reaction energy diagram for the reaction of ethylene with HBr. Two separate steps are involved, each with its own transition state. The energy minimum between the two steps represents the carbocation reaction intermediate.

Each step in a multistep process can always be considered separately. Each step has its own  $\Delta G^\ddagger$  (rate) and its own  $\Delta G^\circ$  (energy change). The *overall*  $\Delta G^\circ$  of the reaction, however, is the energy difference between initial reactants (far left) and final products (far right). This is always true regardless of the shape of the reaction energy curve. Figure 5.7 illustrates some different possible cases.

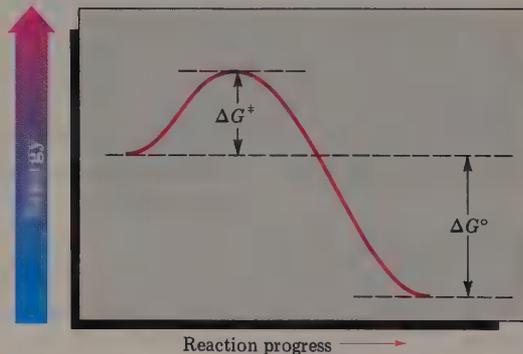
**PRACTICE PROBLEM**.....

Sketch a reaction energy diagram for a one-step reaction that is very fast and highly exothermic.



**Figure 5.7** Hypothetical reaction energy diagrams for some two-step reactions. The overall  $\Delta G^\circ$  for any reaction, regardless of complexity, is simply the energy difference between starting materials and final product. Note that reaction (a) is exothermic, whereas reaction (b) is endothermic.

**Solution** A very fast reaction has a low  $\Delta G^\ddagger$ , and a highly exothermic reaction has a large negative  $\Delta G^\circ$ . Thus, the diagram will look like this:

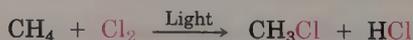


**PROBLEM** .....

- 5.14 Draw a reaction energy diagram for a two-step reaction with an endothermic first step and an exothermic second step. Label the parts of the diagram corresponding to starting material, product, and intermediate.

**PROBLEM** .....

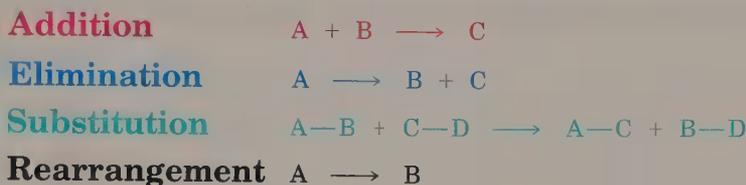
- 5.15 Draw a reaction energy diagram showing both propagation steps in the radical reaction of chlorine with methane. Is the overall  $\Delta G^\circ$  for this reaction positive or negative? Label the parts of your diagram corresponding to  $\Delta G^\circ$  and  $\Delta G^\ddagger$ .



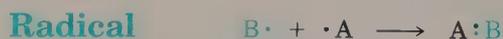
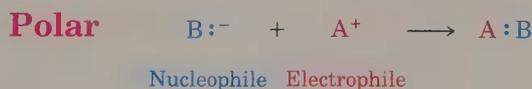
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## 5.11 Summary and Key Words

There are four important kinds of reactions: additions, eliminations, substitutions, and rearrangements. **Addition reactions** take place when two reactants add together to give a single product; **elimination reactions** take place when one reactant splits apart to give two products; **substitution reactions** take place when two reactants exchange parts to give two new products; and **rearrangement reactions** take place when one reactant undergoes a reorganization of bonds and atoms to give a new product.



A full description of how a reaction occurs is called its **mechanism**. There are two kinds of mechanisms by which reactions can take place: **radical mechanisms** and **polar mechanisms**. Polar reactions, the most common type, occur as the result of attractive interactions between an electron-rich (**nucleophilic**) site in the functional group of one molecule and an electron-poor (**electrophilic**) site in the functional group of another molecule. Bonds are formed in polar reactions when the nucleophile donates an electron pair to the electrophile. This movement of electrons is indicated by a curved arrow showing the direction of electron travel from the nucleophile to the electrophile. Radical reactions involve odd-electron species; bonds are formed when each reactant donates one electron to the new bond.



The energy changes that take place during reactions can be described by considering both rates (how fast reaction occurs) and equilibria (how much reaction occurs). The position of a chemical equilibrium is determined by the magnitude of the **Gibbs free-energy change** ( $\Delta G^\circ$ ) that takes place during reaction. The free-energy change is composed of two parts,  $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ . The **enthalpy** term ( $\Delta H^\circ$ ) corresponds to the net change in strength of chemical bonds broken and formed during reaction. The **entropy** term ( $\Delta S^\circ$ ) corresponds to the change in the amount of disorder during reaction. Since the enthalpy term is usually larger and more important than the entropy term, chemists often make the assumption that  $\Delta G^\circ \approx \Delta H^\circ$ .

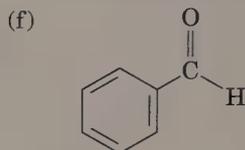
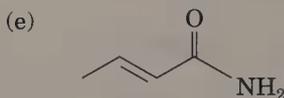
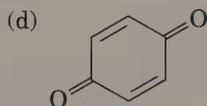
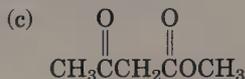
Reactions can be described pictorially using **reaction energy diagrams** that follow the reaction course from starting material through transition state to product. The **transition state** is an activated complex occurring at the highest-energy point during reaction. The amount of energy needed by reactants to reach this high point is the **activation energy**,  $\Delta G^\ddagger$ . It is the

magnitude of  $\Delta G^\ddagger$  that determines the rate of the reaction: The higher the activation energy, the slower the reaction. Most organic reactions have activation energies in the range of 10–35 kcal/mol.

Many reactions take place in more than one step and involve the formation of **reaction intermediates**. An intermediate is a species that lies at an energy minimum between steps on the reaction curve and is formed briefly during the course of a reaction.

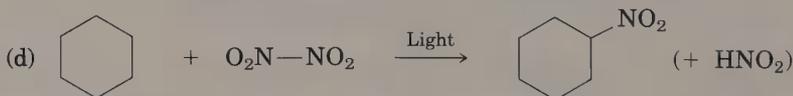
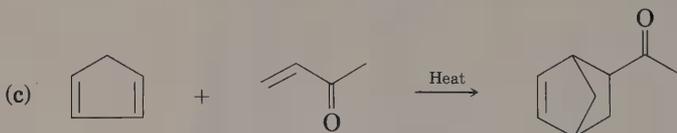
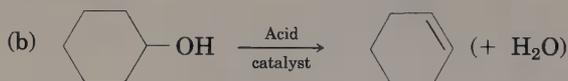
### ADDITIONAL PROBLEMS .....

5.16 Identify the functional groups in these molecules:



5.17 Predict the direction of polarity of the functional groups you identified in Problem 5.16.

5.18 Identify these reactions as being additions, eliminations, substitutions, or rearrangements:



5.19 Explain the differences between addition, elimination, substitution, and rearrangement reactions.

5.20 Define the following:

(a) Polar reaction

(b) Heterolytic bond breakage

(c) Homolytic bond breakage

(d) Radical reaction

(e) Functional group

(f) Polarization

5.21 Give an example of each of the following:

(a) A nucleophile

(b) An electrophile

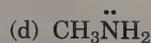
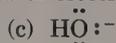
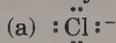
(c) A polar reaction

(d) A substitution reaction

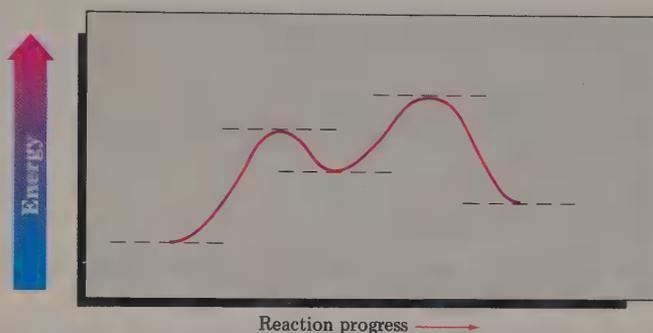
(e) A heterolytic bond breakage

(f) A homolytic bond breakage

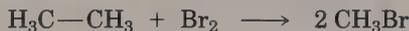
5.22 Classify the following as either nucleophiles or electrophiles.



- 5.23 Draw a reaction energy diagram for a one-step endothermic reaction. Label the parts of the diagram corresponding to reactants, products, transition state,  $\Delta G^\circ$ , and  $\Delta G^\ddagger$ . Is  $\Delta G^\circ$  positive or negative?
- 5.24 Draw a reaction energy diagram for a two-step exothermic reaction. Label the overall  $\Delta G^\circ$ , transition states, and intermediate. Is  $\Delta G^\circ$  positive or negative?
- 5.25 Describe the difference between a transition state and an intermediate.
- 5.26 Draw a reaction energy diagram for a two-step exothermic reaction whose second step is faster than its first step.
- 5.27 Draw a reaction energy diagram for a reaction with  $K_{\text{eq}} = 1$ . What is the value of  $\Delta G^\circ$  in this reaction?
- 5.28 Look at the reaction energy diagram shown here and answer the following questions:



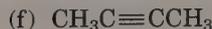
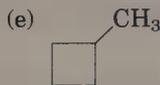
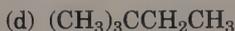
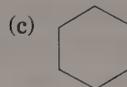
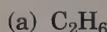
- (a) Indicate  $\Delta G^\circ$  for the reaction. Is it positive or negative?
- (b) How many steps are involved in the reaction?
- (c) Which step is faster?
- (d) How many transition states are there? Label them.
- 5.29 Use the information in Table 5.4 to calculate  $\Delta H^\circ$  for these reactions:
- (a)  $\text{CH}_3\text{OH} + \text{HBr} \longrightarrow \text{CH}_3\text{Br} + \text{H}_2\text{O}$
- (b)  $\text{CH}_3\text{CH}_2\text{OH} + \text{CH}_3\text{Cl} \longrightarrow \text{CH}_3\text{CH}_2\text{OCH}_3 + \text{HCl}$
- 5.30 Use the information in Table 5.4 to calculate  $\Delta H^\circ$  for the reaction of ethane with chlorine, bromine, and iodine:
- (a)  $\text{CH}_3\text{CH}_3 + \text{Cl}_2 \longrightarrow \text{CH}_3\text{CH}_2\text{Cl} + \text{HCl}$
- (b)  $\text{CH}_3\text{CH}_3 + \text{Br}_2 \longrightarrow \text{CH}_3\text{CH}_2\text{Br} + \text{HBr}$
- (c)  $\text{CH}_3\text{CH}_3 + \text{I}_2 \longrightarrow \text{CH}_3\text{CH}_2\text{I} + \text{HI}$
- What can you conclude about the relative ease of chlorination, bromination, and iodination?
- 5.31 An alternative course for the reaction of bromine with ethane could result in the formation of bromomethane:



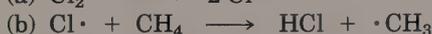
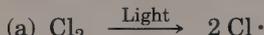
Calculate  $\Delta H^\circ$  for this reaction, and comment on how it compares with the value you calculated in Problem 5.30 for the formation of bromoethane.

- 5.32 Radical chlorination of alkanes is of interest mechanistically but is of little general utility because mixtures of products usually result when more than one kind of C-H bond is present in the substrate. Calculate approximate  $\Delta H^\circ$  values for the possible monochlorination reactions of 2-methylbutane. Use the bond dissociation energies measured for  $\text{CH}_3\text{CH}_2-\text{H}$ ,  $\text{H}-\text{CH}(\text{CH}_3)_2$ , and  $\text{H}-\text{C}(\text{CH}_3)_3$  as representative of typical primary, secondary, and tertiary C-H bonds.

- 5.33 Give the IUPAC name for each of the products formed in Problem 5.32.
- 5.34 Despite the limitations of radical halogenation of hydrocarbons, the reaction is still useful for synthesizing certain halogenated compounds. For which of the following compounds does radical halogenation give a single monohalogenation product?



- 5.35 We've said that the chlorination of methane proceeds by the following steps:

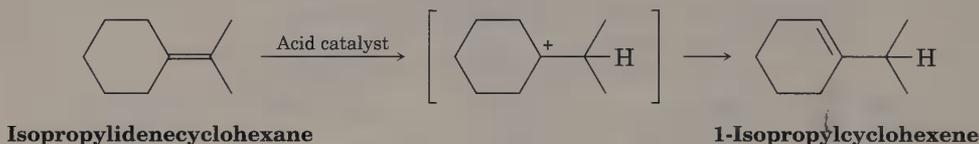


Alternatively, one might propose a different series of steps:



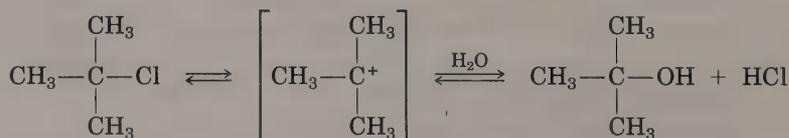
Calculate  $\Delta H^\circ$  for each step in both possible routes. What insight does this provide into the relative merits of each route?

- 5.36 When isopropylidenecyclohexane is treated with strong acid at room temperature, isomerization occurs by the mechanism shown below to yield 1-isopropylcyclohexene:



At equilibrium, the product mixture contains about 30% isopropylidenecyclohexane and about 70% 1-isopropylcyclohexene.

- (a) What kind of reaction is occurring? Is the mechanism polar or radical?
- (b) Draw curved arrows to indicate electron flow in each step.
- (c) Calculate  $K_{eq}$  for the reaction.
- (d) Use Table 5.4 to calculate  $\Delta H^\circ$  for the reaction.
- (e) Since the reaction occurs slowly at room temperature, what is its approximate  $\Delta G^\ddagger$ ?
- (f) Draw a quantitative energy diagram for the reaction.
- 5.37 2-Chloro-2-methylpropane reacts with water in a two-step process to yield 2-methyl-2-propanol. The first step is slower than the second, the reaction takes place slowly at room temperature, and the equilibrium constant is near 1.



2-Chloro-2-methylpropane

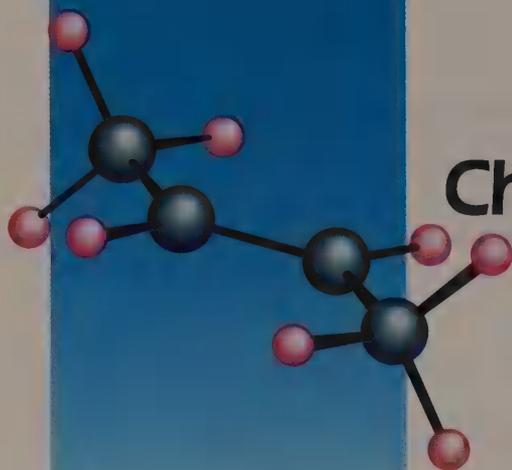
2-Methyl-2-propanol

- (a) Give approximate values for the  $\Delta G^\ddagger$  and  $\Delta G^\circ$  that are consistent with the above information.
- (b) Draw a reaction energy diagram, labeling all points of interest and making sure that the relative energy levels on the diagram correspond to the known data.
- 5.38** The reaction of hydroxide ion with chloromethane to yield methanol and chloride ion is an example of a general reaction type called a nucleophilic substitution reaction.



The value of  $\Delta H^\circ$  for the reaction is  $-18$  kcal/mol and the value of  $\Delta S^\circ$  is  $+13$  cal/mol  $\cdot$  K. What is the value of  $\Delta G^\circ$  at  $25^\circ\text{C}$  (298 K)? Is the reaction exothermic or endothermic?

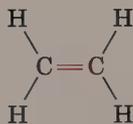
- 5.39** Use the value of  $\Delta G^\circ$  you calculated in Problem 5.38 to find the equilibrium constant  $K_{\text{eq}}$  for the reaction of hydroxide ion with chloride ion.



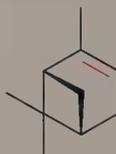
## Chapter 6

# Alkenes: Structure and Reactivity

**Alkenes** are hydrocarbons that contain a carbon–carbon double bond functional group. The word *olefin* is often used as a synonym in the chemical literature, but alkene is the generally preferred term. Alkenes occur abundantly in nature, and many have important biological roles. For example, ethylene is a plant hormone that induces ripening in fruit, and  $\alpha$ -pinene is the major component of turpentine.

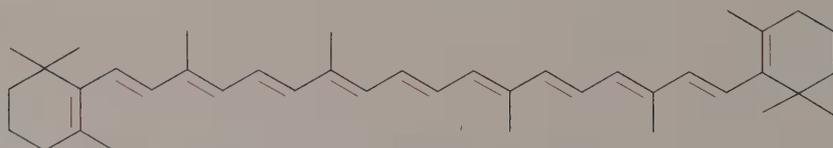


Ethylene



$\alpha$ -Pinene

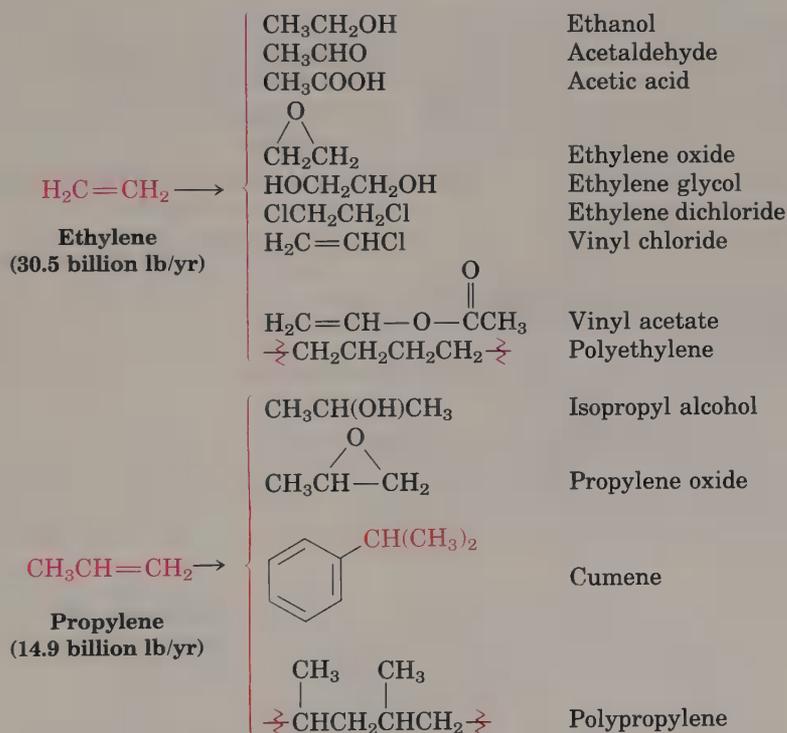
Life itself would be impossible without such alkenes as  $\beta$ -carotene, a compound that contains 11 double bonds.  $\beta$ -Carotene, the orange pigment responsible for the color of carrots, serves as a valuable dietary source of vitamin A and is thought to offer some protection against certain types of cancer.



$\beta$ -Carotene  
(orange pigment and vitamin A precursor)

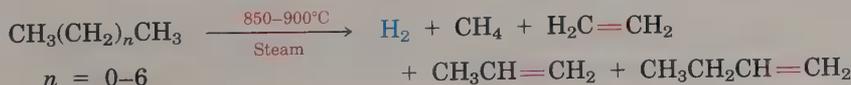
## 6.1 Industrial Preparation and Use of Alkenes

Ethylene (ethene) and propylene (propene), the simplest alkenes, are the two most important organic chemicals produced industrially. More than 35 billion pounds of ethylene and 20 billion pounds of propylene are produced each year in the United States for use in the synthesis of polyethylene, polypropylene, ethylene glycol, acetic acid, acetaldehyde, and a host of other raw materials (Figure 6.1).

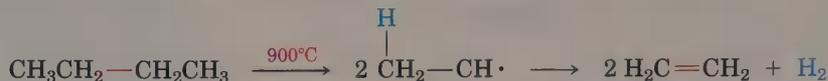


**Figure 6.1** Compounds derived industrially from ethylene (ethene) and propylene (propene).

Ethylene, propylene, and butene are synthesized industrially by thermal cracking of both natural gas ( $\text{C}_1$ – $\text{C}_4$  alkanes) and straight-run gasoline ( $\text{C}_4$ – $\text{C}_8$  *n*-alkanes):



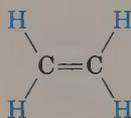
Thermal cracking, introduced in 1912, takes place in the absence of catalysts at temperatures up to 900°C. Although it undoubtedly involves radical reactions, the exact processes are complex. Evidently, the high-temperature reaction conditions cause spontaneous homolysis of carbon–carbon and carbon–hydrogen bonds, with resultant formation of smaller fragments. We might imagine, for instance, that a molecule of butane could split into two ethyl radicals, which could then each lose a hydrogen atom to generate two molecules of ethylene and H<sub>2</sub>:



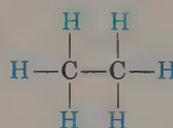
Thermal cracking is an example of a reaction whose energetics are dominated by entropy ( $\Delta S^\circ$ ) rather than by enthalpy ( $\Delta H^\circ$ ) in the free-energy equation  $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ . Although the bond dissociation energy  $\Delta H^\circ$  for a carbon–carbon single bond is high (about 88 kcal/mol), the large positive entropy change resulting from the fragmentation of one large molecule into several smaller pieces, together with the extremely high temperature,  $T$ , makes the  $T\Delta S^\circ$  term larger than the  $\Delta H^\circ$  term. As a result, the cracking reaction is favored.

## 6.2 Calculating the Degree of Unsaturation

Because of its double bond, an alkene has fewer hydrogens than an alkane with the same number of carbons— $\text{C}_n\text{H}_{2n}$  for an alkene versus  $\text{C}_n\text{H}_{2n+2}$  for an alkane—and is therefore referred to as **unsaturated**. Ethylene, for instance, has the formula  $\text{C}_2\text{H}_4$ , whereas ethane has the formula  $\text{C}_2\text{H}_6$ :



Ethylene:  $\text{C}_2\text{H}_4$   
(fewer hydrogens—*unsaturated*)



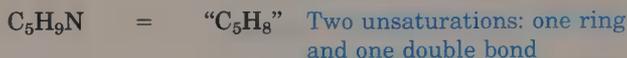
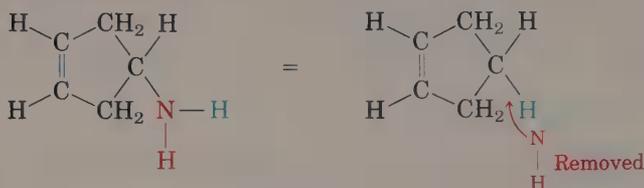
Ethane:  $\text{C}_2\text{H}_6$   
(more hydrogens—*saturated*)

In general, each ring or double bond in a molecule causes a *pair* of hydrogens to be removed from the alkane formula  $\text{C}_n\text{H}_{2n+2}$ . Knowing this relationship, it's possible to work backward from a molecular formula to calculate the **degree of unsaturation** of a molecule—the number of rings and/or multiple bonds present in an unknown.

Let's assume that we want to find the structure of an unknown hydrocarbon. A molecular weight determination on the unknown yields a value of 82, which corresponds to a molecular formula of  $\text{C}_6\text{H}_{10}$ . Since the saturated  $\text{C}_6$  alkane hexane has the formula  $\text{C}_6\text{H}_{14}$ , we can calculate that the unknown compound has two fewer pairs of hydrogens ( $\text{H}_{14} - \text{H}_{10} = \text{H}_4 = 2 \text{H}_2$ ) and its degree of unsaturation is two. The unknown therefore contains



nitrogens from the number of hydrogens to arrive at the equivalent hydrocarbon formula. Again, the best way to convince yourself of this is to see what happens when a nitrogen atom is inserted into an alkane bond: Another hydrogen atom is required to fill the third valency on nitrogen, and we must therefore mentally subtract this extra hydrogen atom to arrive at the equivalent hydrocarbon formula. For example, the formula  $C_5H_9N$  is equivalent to  $C_5H_8$  and thus has two degrees of unsaturation:



PROBLEM .....

6.1 Calculate the degree of unsaturation in these hydrocarbons:

- (a)  $C_8H_{14}$  (b)  $C_5H_6$  (c)  $C_{12}H_{20}$   
 (d)  $C_{20}H_{32}$  (e)  $C_{40}H_{56}$ ,  $\beta$ -carotene

PROBLEM .....

6.2 Calculate the degree of unsaturation in these formulas, and then draw as many structures as you can for each.

- (a)  $C_4H_8$  (b)  $C_4H_6$  (c)  $C_3H_4$

PROBLEM .....

6.3 Calculate the degree of unsaturation in these formulas:

- (a)  $C_6H_5N$  (b)  $C_6H_5NO_2$  (c)  $C_8H_9Cl_3$   
 (d)  $C_9H_{16}Br_2$  (e)  $C_{10}H_{12}N_2O_3$  (f)  $C_{20}H_{32}O_2$

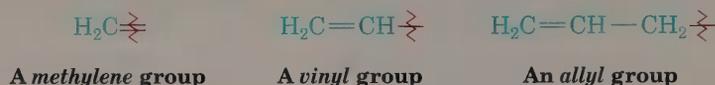
## 6.3 Nomenclature of Alkenes

Alkenes are named according to a series of rules similar to those developed for alkanes, with the suffix *-ene* used instead of *-ane* to identify the family. There are three steps:

- Step 1* Name the parent hydrocarbon. Find the longest carbon chain containing the double bond, and name the compound accordingly, using the suffix *-ene*:



For historical reasons, there are a small number of alkenes whose names, though firmly entrenched in common usage, don't conform to strict rules of nomenclature. For example, the alkene derived from ethane should properly be called *ethene*, but the name *ethylene* has been used so long that it is accepted by IUPAC. Table 6.1 lists several other common names that are often used and that are recognized by IUPAC. Note that a  $=\text{CH}_2$  substituent is called a **methylene group**, a  $\text{H}_2\text{C}=\text{CH}-$  substituent is called a **vinyl group**, and a  $\text{H}_2\text{C}=\text{CH}-\text{CH}_2-$  substituent is called an **allyl group**:

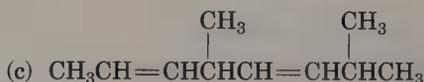
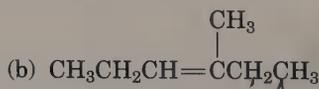
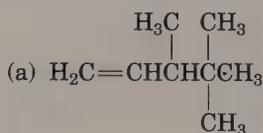
TABLE 6.1 Common Names of Some Alkenes<sup>a</sup>

Compound	Systematic name	Common name
$\text{H}_2\text{C}=\text{CH}_2$	Ethene	Ethylene
$\text{CH}_3\text{CH}=\text{CH}_2$	Propene	Propylene
$\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3\text{C}=\text{CH}_2 \end{array}$	2-Methylpropene	Isobutylene
$\begin{array}{c} \text{CH}_3 \\   \\ \text{H}_2\text{C}=\text{C}-\text{CH}=\text{CH}_2 \end{array}$	2-Methyl-1,3-butadiene	Isoprene
$\text{CH}_3\text{CH}=\text{CHCH}=\text{CH}_2$	1,3-Pentadiene	Piperylene
$\text{H}_2\text{C}=\text{CH} \begin{array}{c} \nearrow \\ \searrow \end{array}$	Ethenyl	Vinyl (an alkenyl group)
$\text{H}_2\text{C}=\text{CH}-\text{CH}_2 \begin{array}{c} \nearrow \\ \searrow \end{array}$	2-Propenyl	Allyl
$\text{H}_2\text{C} \begin{array}{c} \nearrow \\ \searrow \end{array}$	Methylene	
$\text{CH}_3\text{CH} \begin{array}{c} \nearrow \\ \searrow \end{array}$	Ethylidene	

<sup>a</sup>Both common and systematic names are recognized by IUPAC.

## PROBLEM .....

6.4 Give proper IUPAC names for these compounds:



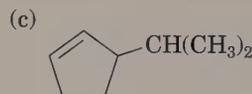
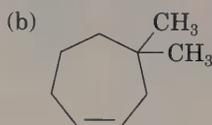
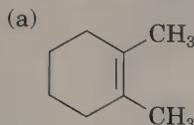
## PROBLEM .....

6.5 Draw structures corresponding to these IUPAC names:

- (a) 2-Methyl-1,5-hexadiene      (b) 3-Ethyl-2,2-dimethyl-3-heptene  
 (c) 2,3,3-Trimethyl-1,4,6-octatriene      (d) 3,4-Diisopropyl-2,5-dimethyl-3-hexene  
 (e) 4-*tert*-Butyl-2-methylheptane

## PROBLEM .....

6.6 Name these cycloalkenes:



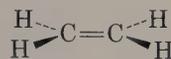
## 6.4 Electronic Structure of Alkenes

We saw in Section 1.10 that the carbon atoms in a double bond are  $sp^2$  hybridized and have three equivalent orbitals that lie in a plane at angles of  $120^\circ$  to one another. The fourth carbon orbital is an unhybridized  $p$  orbital, which is perpendicular to the  $sp^2$  plane. When two such carbon atoms approach each other, they form two kinds of bonds: a sigma bond, formed by head-on overlap of  $sp^2$  orbitals, and a pi bond, formed by sideways overlap of  $p$  orbitals. As is always true when orbitals overlap, both bonding and antibonding molecular orbitals result (Section 1.7). Only the bonding combinations are filled, however; the higher-energy antibonding orbitals remain vacant (Figure 6.2).

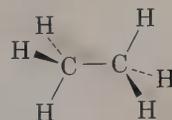
As you might expect, a carbon-carbon double bond is both stronger (152 kcal/mol versus 88 kcal/mol; 636 kJ/mol versus 368 kJ/mol) and shorter (1.33 Å versus 1.54 Å) than a carbon-carbon single bond. Table 6.2 compares the experimentally determined bond parameters of ethylene and ethane. The doubly bonded carbons and the four atoms attached to them lie in a plane, with bond angles of approximately  $120^\circ$ .

TABLE 6.2 Molecular Parameters for Ethylene and Ethane<sup>a</sup>

	<i>Ethylene</i>	<i>Ethane</i>
H—C—H bond angle (degrees)	116.6	109.3
H—C—C bond angle (degrees)	121.7	109.6
C—C bond strength (kcal/mol)	152	88
C—C bond length (Å)	1.33	1.54
C—H bond strength (kcal/mol)	107	98
C—H bond length (Å)	1.076	1.10

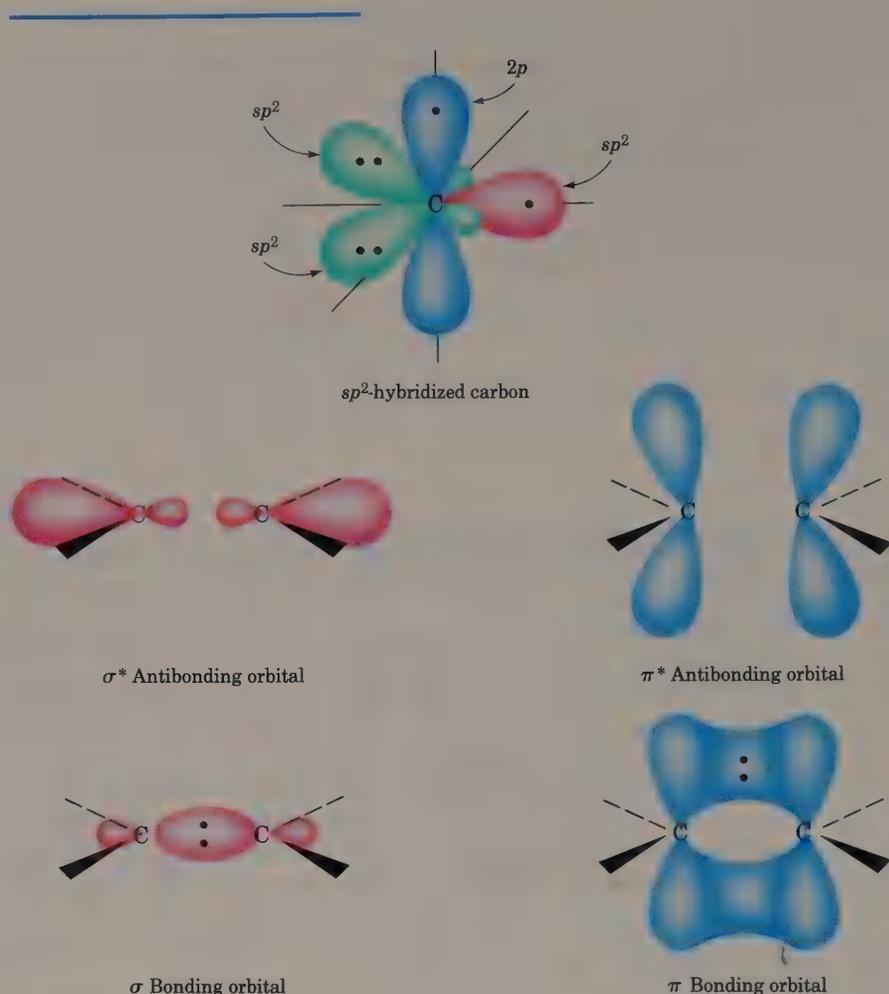


Ethylene



Ethane

<sup>a</sup>The double bond is both stronger and shorter than the single bond.

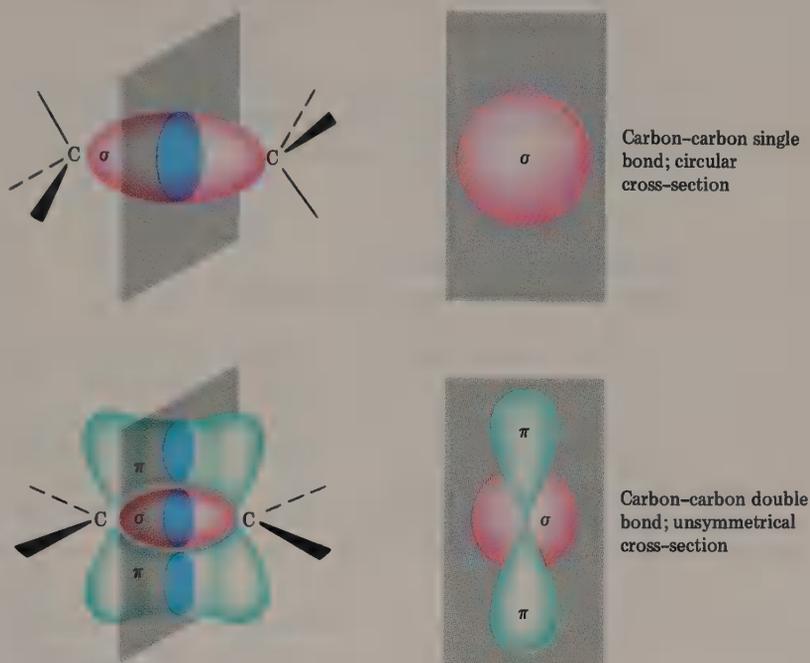


**Figure 6.2** An orbital picture of the carbon–carbon double bond. Both sigma and pi bonding molecular orbitals are filled. The antibonding combinations are unfilled.

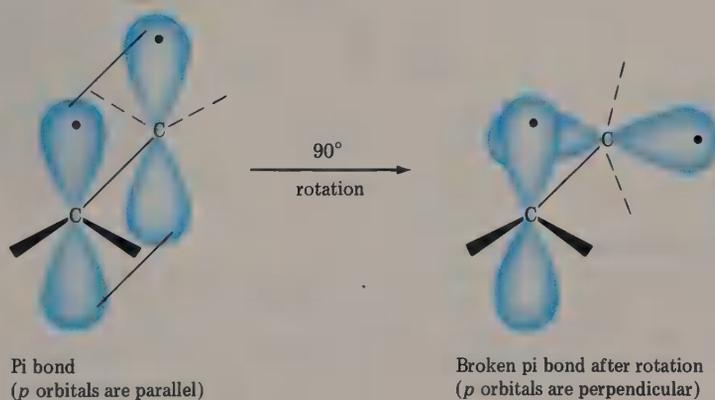
We know from Section 4.1 that relatively free rotation is possible around sigma bonds and that open-chain alkanes like butane therefore have many rapidly interconverting conformations. The same is not true for double bonds, however: Carbon–carbon double bonds do not have circular cross-sections, and rotation is therefore restricted (Figure 6.3).

If we were to *force* rotation to occur around a double bond, we would need to break the pi bond temporarily (Figure 6.4). Thus, the barrier to double-bond rotation must be at least as great as the strength of the pi bond itself.

We can make a rough estimate of how much energy is required to break the pi bond of an alkene by subtracting the value for the strength of an average carbon–carbon sigma bond (88 kcal/mol) from the total bond strength value for ethylene (152 kcal/mol). This calculation predicts an



**Figure 6.3** Cross sections cut through carbon-carbon single and double bonds. Free rotation is possible around a single bond but not around a double bond.



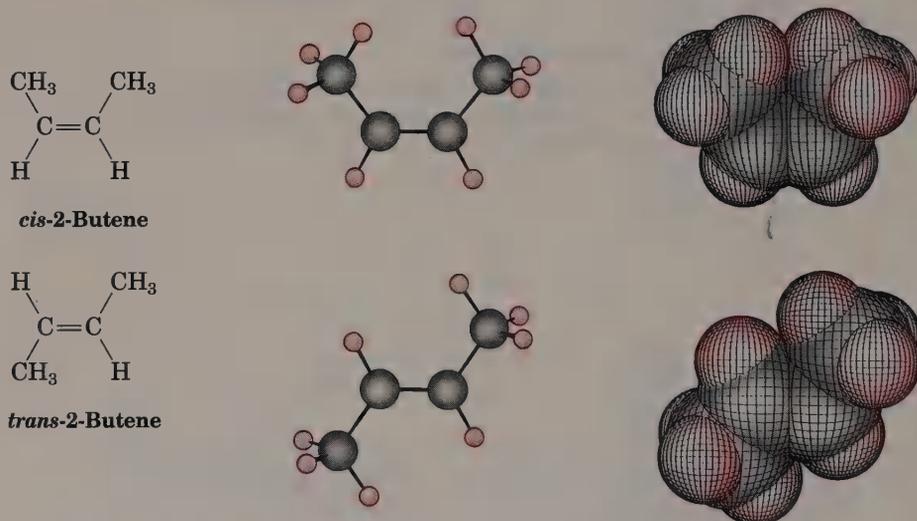
**Figure 6.4** The pi bond must break in order for rotation to take place around a carbon-carbon double bond.

approximate bond strength of 64 kcal/mol (268 kJ/mol) for the ethylene pi bond, and it's therefore clear why rotation can't occur. (Recall that the barrier to bond rotation in ethane is only 2.9 kcal/mol.)

Ethylene C=C bond strength (sigma + pi)	152 kcal/mol
Ethane C—C bond strength (sigma only)	<u>88 kcal/mol</u>
Difference (pi bond only)	64 kcal/mol

## 6.5 Cis–Trans Isomerism in Alkenes

The lack of rotation around the carbon–carbon double bond is of more than just theoretical interest; it also has chemical consequences. Imagine the situation for a disubstituted alkene such as 2-butene. (*Disubstituted* means that two substituents other than hydrogen are bonded to the double-bond carbons.) The two methyl groups in 2-butene can either be on the same side of the double bond or on opposite sides, a situation reminiscent of substituted cycloalkanes (Section 3.9). Figure 6.5 shows the two 2-butene isomers.

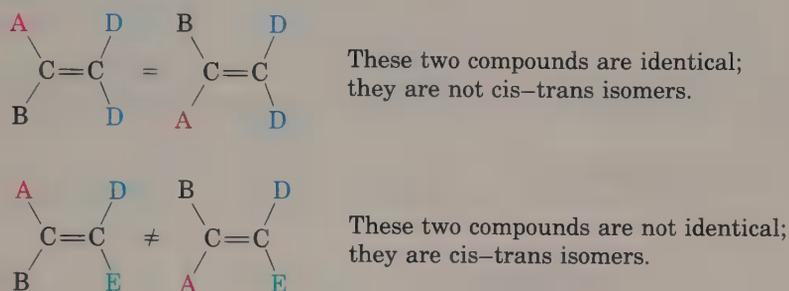


**Figure 6.5** Cis and trans isomers of 2-butene. The cis isomer has the substituent methyl groups on the same side of the double bond, and the trans isomer has the methyl groups on opposite sides.

Since bond rotation can't occur, the two 2-butenes can't spontaneously interconvert; they are distinct, isolable compounds. As with disubstituted cycloalkanes (Section 3.9), we call such compounds **cis–trans stereoisomers** because they have the same formula and overall skeleton but differ

in the spatial arrangement of their atoms. The compound with substituents on the same side of the double bond is referred to as *cis*-2-butene; the isomer with substituents on opposite sides is *trans*-2-butene.

*Cis-trans* isomerism is not limited to disubstituted alkenes. It can occur whenever both of the double-bond carbons are attached to two different groups. If one of the double-bond carbons is attached to two identical groups, however, then *cis-trans* isomerism is not possible (Figure 6.6).



**Figure 6.6** The requirement for *cis-trans* isomerism in alkenes. Compounds that have one of their carbons bonded to two identical groups can't exist as *cis-trans* isomers.

PROBLEM .....

- 6.7 Which of the following compounds can exist as pairs of *cis-trans* isomers? Draw each *cis-trans* pair, and indicate the geometry of each isomer.
- |   |   |
|---|---|
| (a) $\text{CH}_3\text{CH}=\text{CH}_2$              | (b) $(\text{CH}_3)_2\text{C}=\text{CHCH}_3$                               |
| (c) $\text{CH}_3\text{CH}_2\text{CH}=\text{CHCH}_3$ | (d) $(\text{CH}_3)_2\text{C}=\text{C}(\text{CH}_3)\text{CH}_2\text{CH}_3$ |
| (e) $\text{ClCH}=\text{CHCl}$                       | (f) $\text{BrCH}=\text{CHCl}$   |

PROBLEM .....

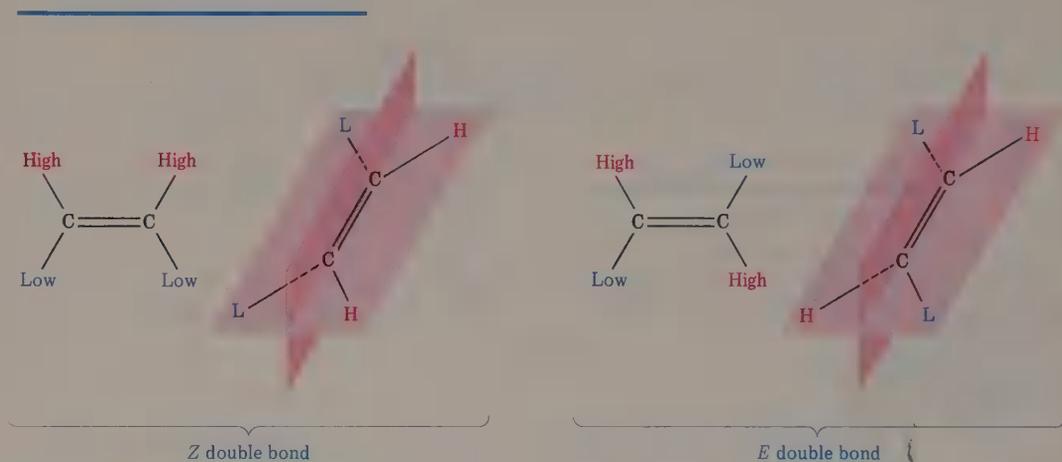
- 6.8 How can you account for the observation that cyclohexene does not show *cis-trans* double-bond isomerism, whereas cyclodecene can exist in both *cis* and *trans* forms? Making molecular models should be helpful.
- .....

## 6.6 Sequence Rules: The *E,Z* Designation

In the previous discussion of isomerism in the 2-butenes, we used the terms *cis* and *trans* to denote alkenes whose substituents were on the same side and opposite side of a double bond, respectively. This *cis-trans* nomenclature is unambiguous for all disubstituted alkenes, but how do we denote

the geometry of *trisubstituted* and *tetrasubstituted* double bonds? (*Trisubstituted* means three substituents other than hydrogen attached to the double-bond carbons; *tetrasubstituted* means four substituents other than hydrogen.)

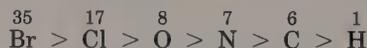
The answer is provided by the *E,Z* system of nomenclature, which uses a series of **sequence rules** to assign priorities to the substituent groups on the double-bond carbons. Considering each of the double-bond carbons separately, we use the sequence rules to decide which of the two attached groups is higher in priority. If the higher-priority groups on each carbon are on the same side of the double bond, the alkene is designated *Z* (for the German *zusammen*, “together”). If the higher-priority groups are on opposite sides, the alkene is designated *E* (for the German *entgegen*, “opposite”). The easiest way to learn which is which is to think with an accent: *Z* = groups on “ze same zide” (*E* = the other one). These assignments are shown in Figure 6.7.



**Figure 6.7** The *E,Z* system of nomenclature for substituted alkenes.

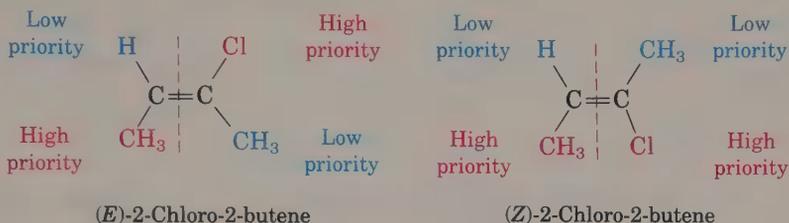
Introduced by Cahn, Ingold, and Prelog<sup>1</sup> in 1964, the sequence rules are as follows:

**Sequence rule 1** Look at the atoms directly attached to each of the double-bond carbons and rank them in order of decreasing atomic number. That is, an atom with higher atomic number like Cl (17) receives higher priority than an atom with lower atomic number like H (1). Thus, the common atoms that we might find attached to a double bond would be assigned the following priority sequence:



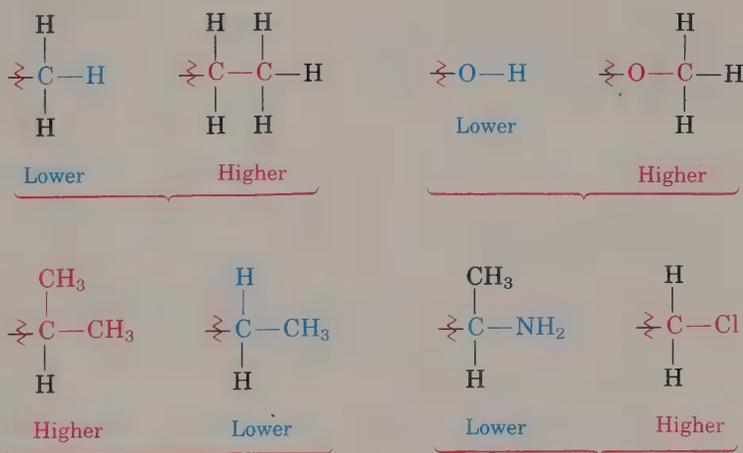
<sup>1</sup>Vladimir Prelog (1906– ); b. Sarajevo (now in Yugoslavia); Dr. Ing., Institute of Technology, Prague (Votocék); Professor, University of Zagreb, Federal Institute of Technology (ETH), Zürich (1941–1976); Nobel Prize (1975).

For example:

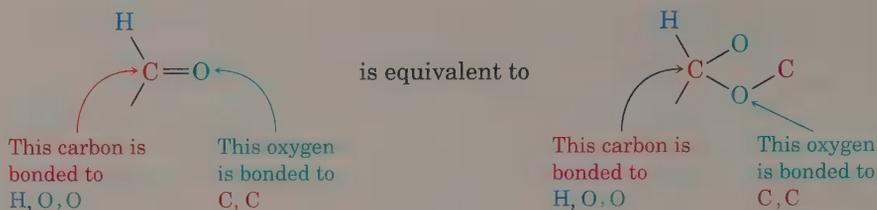


Since chlorine has a higher atomic number than carbon (17 versus 6), it receives higher priority than a methyl (CH<sub>3</sub>) group. Methyl receives higher priority than hydrogen, however, and the left-hand isomer in the preceding display is therefore assigned *E* geometry (high-priority groups on opposite sides of the double bond). The right-hand isomer has *Z* geometry (high-priority groups on ze zame zide of the double bond).

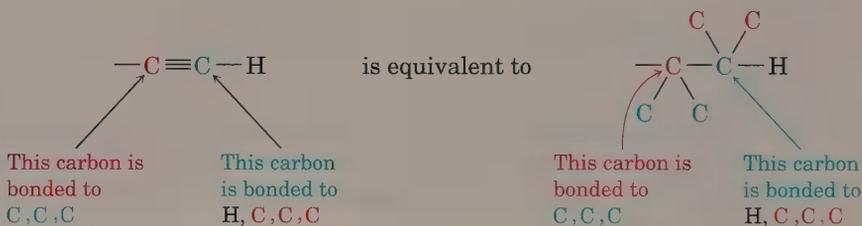
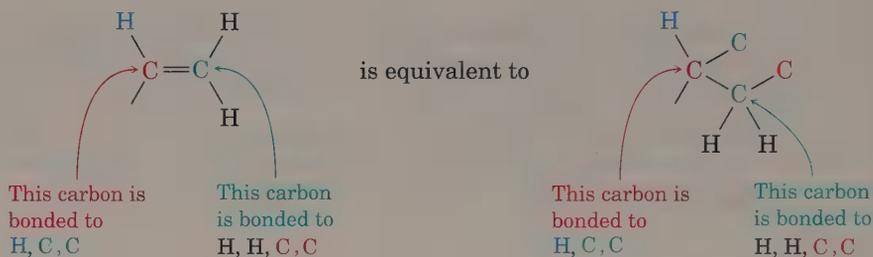
**Sequence rule 2** If a decision can't be reached by ranking the first atoms in the substituent, look at the second, third, or fourth atoms away from the double-bond carbons until a difference is found. Thus, an ethyl substituent, -CH<sub>2</sub>CH<sub>3</sub>, and a methyl substituent, -CH<sub>3</sub>, are equivalent by rule 1 since both have carbon as the first atom. By rule 2, however, ethyl receives higher priority than methyl because its *second* atoms are one carbon and two hydrogens rather than three hydrogens. Look at the following examples to see how the rule works:



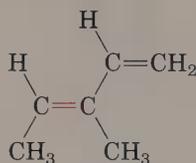
**Sequence rule 3** Multiple-bonded atoms are considered equivalent to the same number of single-bonded atoms. For example, an aldehyde substituent (-CH=O), which has a carbon atom *doubly* bonded to *one* oxygen, is equivalent to a substituent with a carbon atom *singly* bonded to *two* oxygen atoms:



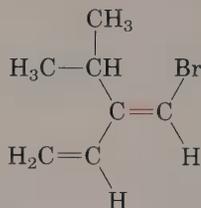
As further examples, the following pairs are equivalent:



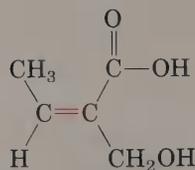
Taking all the sequence rules into account, we can assign the configurations shown in the following examples. Work through each one to convince yourself that the assignments are correct.



(*E*)-3-Methyl-1,3-pentadiene



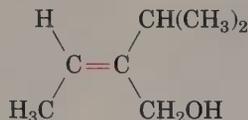
(*E*)-1-Bromo-2-isopropyl-1,3-butadiene



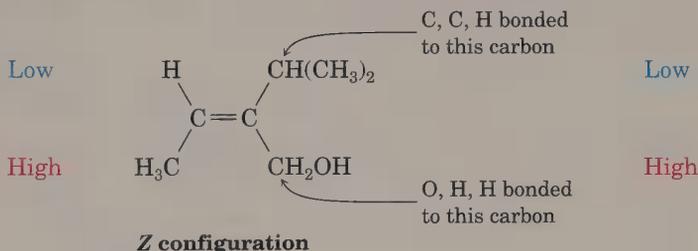
(*Z*)-2-Hydroxymethyl-2-butenic acid

## PRACTICE PROBLEM.....

Assign *E* or *Z* configuration to the double bond in this compound:



**Solution** Look at the two double-bond carbons individually. The left-hand carbon has two substituents,  $-\text{H}$  and  $-\text{CH}_3$ , of which  $-\text{CH}_3$  receives higher priority by sequence rule 1. The right-hand carbon also has two substituents,  $-\text{CH}(\text{CH}_3)_2$  and  $-\text{CH}_2\text{OH}$ , but rule 1 doesn't allow a priority assignment to be made since both groups have carbon as their first atom. By rule 2, however,  $-\text{CH}_2\text{OH}$  receives higher priority than  $-\text{CH}(\text{CH}_3)_2$ , since  $-\text{CH}_2\text{OH}$  has an *oxygen* and two hydrogens as its second atoms, whereas  $-\text{CH}(\text{CH}_3)_2$  has two *carbons* and one hydrogen as its second atoms. Thus, the two high-priority groups are on the same side of the double bond, and we assign *Z* configuration.



## PROBLEM.....

6.9 Which member in each set is higher in priority?

- (a)  $-\text{H}$  or  $-\text{Br}$                       (b)  $-\text{Cl}$  or  $-\text{Br}$                       (c)  $-\text{CH}_3$  or  $-\text{CH}_2\text{CH}_3$   
 (d)  $-\text{NH}_2$  or  $-\text{OH}$                       (e)  $-\text{CH}_2\text{OH}$  or  $-\text{CH}_3$                       (f)  $-\text{CH}_2\text{OH}$  or  $-\text{CH}=\text{O}$

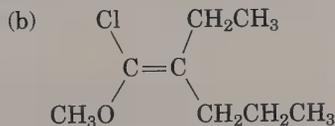
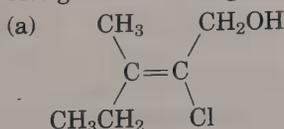
## PROBLEM.....

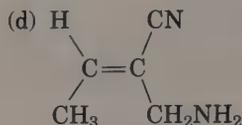
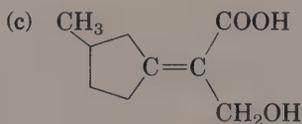
6.10 Rank the sets of substituents in order of Cahn-Ingold-Prelog priorities:

- (a)  $-\text{CH}_3$ ,  $-\text{OH}$ ,  $-\text{H}$ ,  $-\text{Cl}$   
 (b)  $-\text{CH}_3$ ,  $-\text{CH}_2\text{CH}_3$ ,  $-\text{CH}=\text{CH}_2$ ,  $-\text{CH}_2\text{OH}$   
 (c)  $-\text{COOH}$ ,  $-\text{CH}_2\text{OH}$ ,  $-\text{C}\equiv\text{N}$ ,  $-\text{CH}_2\text{NH}_2$   
 (d)  $-\text{CH}_2\text{CH}_3$ ,  $-\text{C}\equiv\text{CH}$ ,  $-\text{C}\equiv\text{N}$ ,  $-\text{CH}_2\text{OCH}_3$

## PROBLEM.....

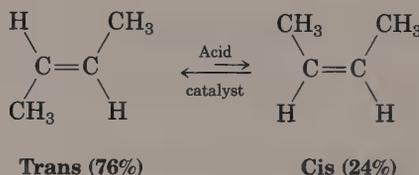
6.11 Assign *E* or *Z* configuration to these alkenes:





## 6.7 Alkene Stability

Although the *cis*-*trans* interconversion of alkene isomers does not occur spontaneously, it can be made to happen under appropriate conditions—for example, on treatment with a strong acid catalyst. If we were to interconvert *cis*-2-butene with *trans*-2-butene and allow them to reach equilibrium, we would find that they aren't of equal stability. At equilibrium, the *trans* isomer is more favored than the *cis* isomer by a ratio of 76% *trans* to 24% *cis*.

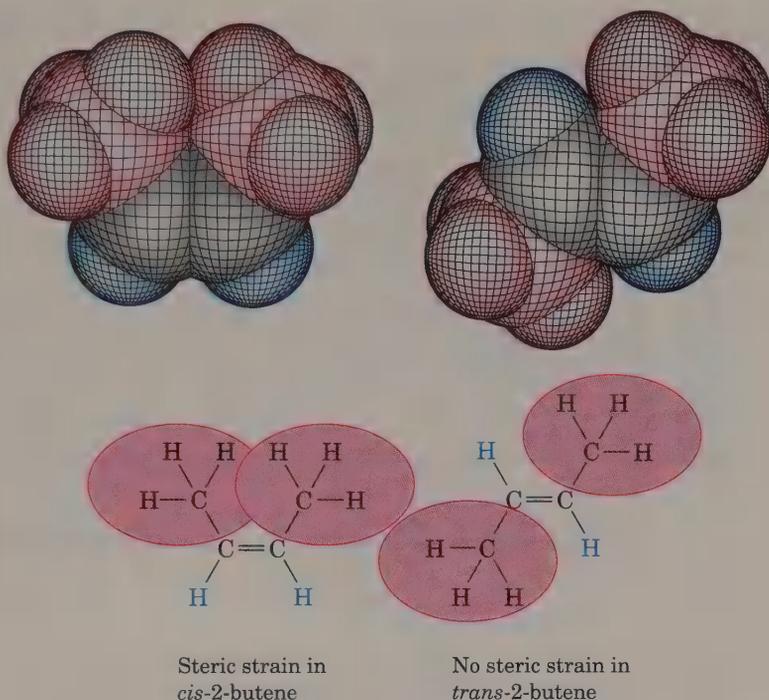


Using the relationship between equilibrium constants and free-energy differences in Table 6.3 (which repeats Table 4.3 for convenience), we can calculate that *cis*-2-butene is less stable than *trans*-2-butene by 0.66 kcal/mol (2.6 kJ/mol) at room temperature.

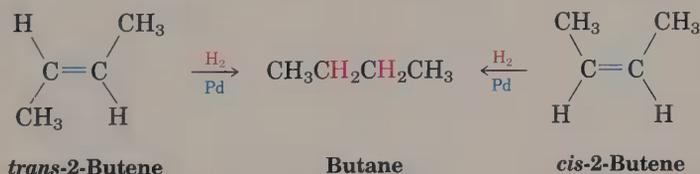
**TABLE 6.3** The Relationship Between Stability and Isomer Percentages at Equilibrium

More stable isomer (%)	Less stable isomer (%)	Energy difference at 25°C	
		(kcal/mol)	(kJ/mol)
50	50	0	0
75	25	0.65	2.72
90	10	1.30	5.45
95	5	1.74	7.29
99	1	2.72	11.38
99.9	0.1	4.09	17.11

*Cis* alkenes are less stable than their *trans* isomers because of steric (spatial) strain between the two bulky substituents on the same side of the double bond. This is the same kind of steric interference that we saw previously in the axial conformation of methylcyclohexane (Section 4.12).

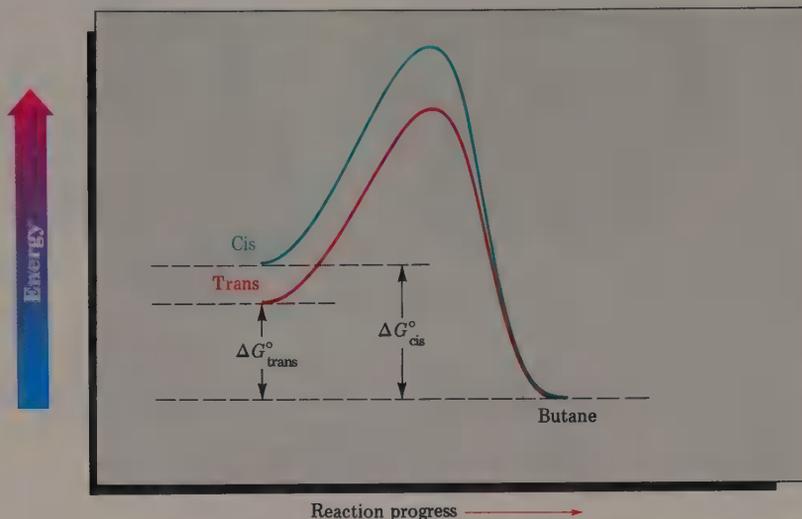


Although it's sometimes possible to obtain information about relative stabilities of alkenes by establishing a *cis*–*trans* equilibrium through treatment of the alkene with strong acid, there's an easier way to gain nearly the same information. One of the more important reactions that alkenes undergo is called *catalytic hydrogenation*. In the presence of a catalyst such as palladium or platinum, hydrogen adds to a carbon–carbon double bond to yield the corresponding saturated alkane:



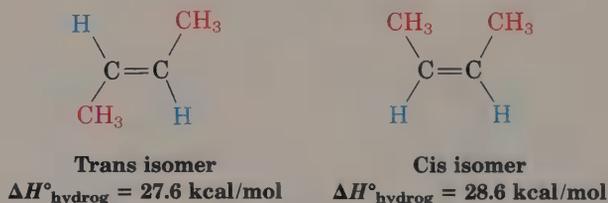
Consider the hydrogenations of *cis*- and *trans*-2-butene, both of which react with hydrogen to give the same product, butane. Energy diagrams for the two reactions are shown in Figure 6.8. Since *cis*-2-butene is less stable than *trans*-2-butene by 0.66 kcal/mol, the energy diagram shows the *cis* alkene at a higher energy level. After reaction, however, both curves are at the same energy level (butane). It therefore follows that  $\Delta G^\circ$  for reaction of the *cis* isomer must be larger than for reaction of the *trans* isomer. In other words, more energy is evolved in the hydrogenation of the *cis* isomer than of the *trans* isomer because there was more energy present in the *cis* isomer to begin with.

If we were to measure the heats of reaction for the two hydrogenations and find their difference, we could determine the relative stabilities of *cis* and *trans* isomers without having to measure an equilibrium position. A large number of such **heats of hydrogenation** ( $\Delta H^\circ_{\text{hydrog}}$ ) have been



**Figure 6.8** Reaction energy diagrams for hydrogenation of *cis*- and *trans*-2-butene. The *cis* isomer is higher in energy than the *trans* isomer by about 0.66 kcal/mol and therefore gives off more energy in the reaction.

measured, and the results bear out our expectation. For *cis*-2-butene,  $\Delta H^\circ_{\text{hydrog}} = 28.6$  kcal/mol (120 kJ/mol); for the *trans* isomer,  $\Delta H^\circ_{\text{hydrog}} = 27.6$  kcal/mol (116 kJ/mol). (Note that these  $\Delta H^\circ$  values should be negative numbers because heat is evolved during hydrogenation. The minus sign is usually dropped, however, to make comparisons easier.)



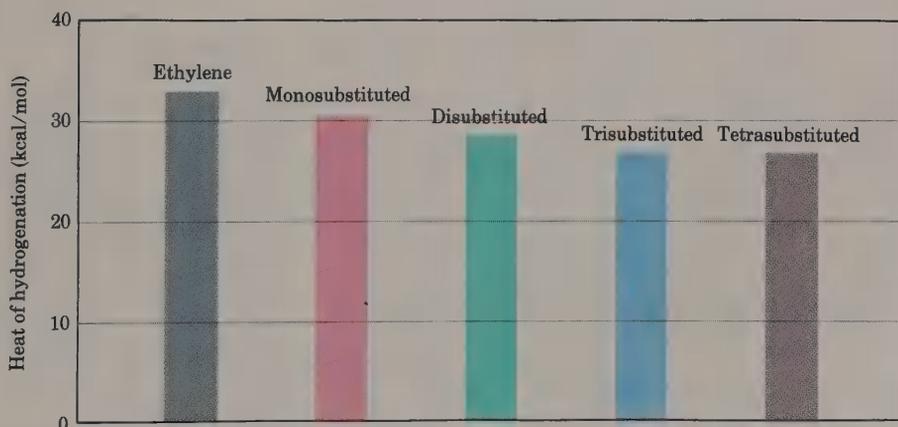
Although the energy difference in the heats of hydrogenation for the 2-butene isomers (1 kcal/mol) agrees reasonably well with the energy difference calculated from equilibrium data (0.66 kcal/mol), the two numbers aren't exactly the same. There are two reasons for the difference. The first is simply experimental error; heats of hydrogenation require considerable expertise and specialized equipment to measure accurately, and we are looking at a small difference between two large numbers. The second reason is that heats of reaction and equilibrium constants don't measure exactly the same quantity. Heats of reaction measure enthalpy changes,  $\Delta H^\circ$ , whereas equilibrium constants measure overall free-energy changes,  $\Delta G^\circ$  ( $\Delta G^\circ = \Delta H^\circ - T\Delta S^\circ$ ). We therefore expect a slight difference when comparing the two measurements.

Although heats of hydrogenation are not as accurate as we might like, we can nevertheless gain some useful and interesting information from them. Table 6.4 lists some representative data, and Figure 6.9 plots the results graphically.

The data in Table 6.4 and in Figure 6.9 show that alkenes become more stable with increasing substitution. For example, ethylene has  $\Delta H^\circ_{\text{hydrog}} = 32.8$  kcal/mol (137 kJ/mol), but when one alkyl substituent is attached as in 1-butene ( $\Delta H^\circ_{\text{hydrog}} = 30.3$  kcal/mol) the alkene becomes

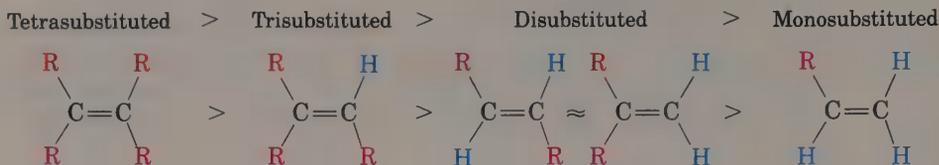
**TABLE 6.4** Heats of Hydrogenation of Some Alkenes

Substitution	Alkene	$\Delta H^\circ_{\text{hydrog}}$	
		(kcal/mol)	(kJ/mol)
	$\text{H}_2\text{C}=\text{CH}_2$	32.8	137
Monosubstituted (one alkyl group next to double bond)	$\text{CH}_3\text{CH}=\text{CH}_2$	30.1	126
	$\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$	30.3	127
	$(\text{CH}_3)_2\text{CHCH}=\text{CH}_2$	30.3	127
	$(\text{CH}_3)_3\text{CCH}=\text{CH}_2$	30.3	127
Disubstituted (two alkyl groups)	Cis $\text{CH}_3\text{CH}=\text{CHCH}_3$	28.6	120
	Trans $\text{CH}_3\text{CH}=\text{CHCH}_3$	27.6	115
	$(\text{CH}_3)_2\text{C}=\text{CH}_2$	28.4	119
	$\text{CH}_3\text{CH}_2(\text{CH}_3)\text{C}=\text{CH}_2$	28.5	119
Trisubstituted (three alkyl groups)	$(\text{CH}_3)_2\text{C}=\text{CHCH}_3$	26.9	113
Tetrasubstituted (four alkyl groups)	$(\text{CH}_3)_2\text{C}=\text{C}(\text{CH}_3)_2$	26.6	111

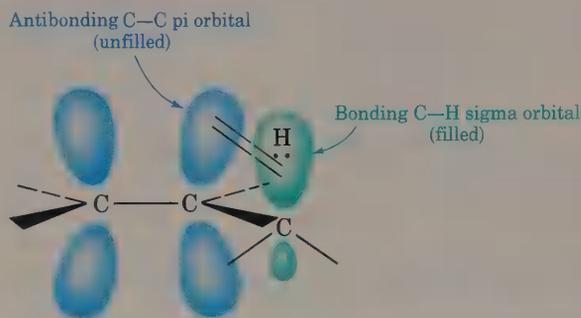


**Figure 6.9** A plot of heat of hydrogenation versus substitution pattern for alkenes. Alkene stability increases with increasing substitution.

approximately 2.5 kcal/mol more stable. Increasing the degree of substitution leads to still further stability. As a general rule, alkenes follow the stability order:

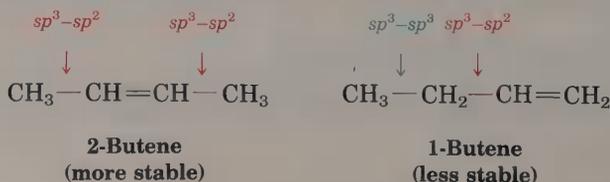


Two explanations have been advanced to account for the observed stability order. Most chemists feel that the stability order is due primarily to **hyperconjugation** (Figure 6.10), a stabilizing effect resulting from overlap between the unfilled antibonding carbon–carbon pi bond orbital and a filled carbon–hydrogen sigma bond orbital on a neighboring substituent. The more substituents that are present, the more opportunities exist for hyperconjugation and the more stable the alkene.



**Figure 6.10** Hyperconjugation—a stabilizing effect due to overlap between an unfilled antibonding pi orbital and a neighboring filled C–H sigma bond orbital.

In addition to the effect of hyperconjugation, a simple bond-strength argument can also be used to explain the observed alkene stability order. A bond between an  $sp^2$  carbon and an  $sp^3$  carbon is somewhat stronger than a bond between two  $sp^3$  carbons. Thus, in comparing 1-butene and 2-butene, the monosubstituted isomer has one  $sp^3$ – $sp^3$  bond and one  $sp^3$ – $sp^2$  bond, and the disubstituted isomer has two  $sp^3$ – $sp^2$  bonds. Highly substituted alkenes always have a higher ratio of  $sp^3$ – $sp^2$  bonds to  $sp^3$ – $sp^3$  bonds than less substituted alkenes and are therefore more stable.



## PROBLEM

- 6.12 Which alkene in each of the following sets is more stable?  
 (a) 1-Butene or 2-methylpropene      (b) (*Z*)-2-Hexene or (*E*)-2-hexene  
 (c) 1-Methylcyclohexene or 3-methylcyclohexene

## PROBLEM

- 6.13 The double bonds in small-ring cycloalkenes must have *cis* geometry, because a stable *trans* double bond is impossible in a five- or six-membered ring. At some point, however, a ring becomes large enough to accommodate a *trans* double bond. The following heats of hydrogenation have been measured:

	$\Delta H^\circ_{\text{hydrog}}$	
	(kcal/mol)	(kJ/mol)
<i>cis</i> -Cyclooctene	23.0	96.2
<i>trans</i> -Cyclooctene	32.2	134.7
<i>cis</i> -Cyclononene	23.6	98.7
<i>trans</i> -Cyclononene	26.5	110.9
<i>cis</i> -Cyclodecene	20.7	86.6
<i>trans</i> -Cyclodecene	24.0	100.4

How do you explain these data? Make molecular models of the *trans* cycloalkenes to see their conformations.

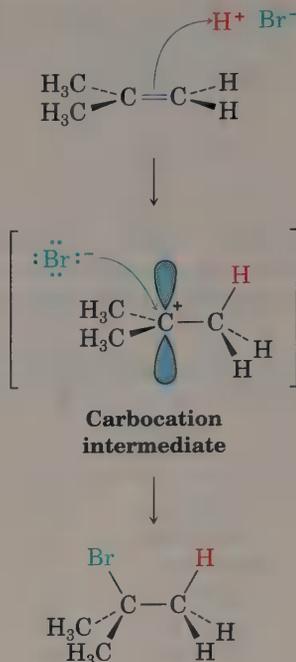
## 6.8 Alkene Electrophilic Addition Reactions

Before beginning a detailed discussion of alkene reactions, let's review briefly some conclusions from the previous chapter (Sections 5.6–5.10). At that time, we said that alkenes behave as nucleophiles. The carbon–carbon double bond is electron-rich and can donate a pair of electrons to an electrophile in polar reactions. For example, reaction of 2-methylpropene with HBr leads to the formation of 2-bromo-2-methylpropane. Careful study of this and similar reactions by Sir Christopher Ingold<sup>2</sup> and others many years ago has led to the generally accepted mechanism shown in Figure 6.11 for what is called an **electrophilic addition reaction**.

The reaction begins with an attack on the electrophile, HBr, by the electrons of the nucleophilic pi bond. Two electrons from the pi bond form a new sigma bond between the entering hydrogen and an alkene carbon, as

<sup>2</sup>Sir Christopher Ingold (1893–1970); b. Ilford, England; D.Sc., London (Thorpe); professor, Leeds (1924–1930), University College, London (1930–1970).

The electrophile  $\text{H}^+$  is attacked by the pi electrons of the double bond, and a new C–H sigma bond is formed. This leaves the other carbon atom with a + charge and a vacant  $p$  orbital.



$\text{Br}^-$  donates an electron pair to the positively charged carbon atom, forming a C–Br sigma bond and yielding the neutral addition product.

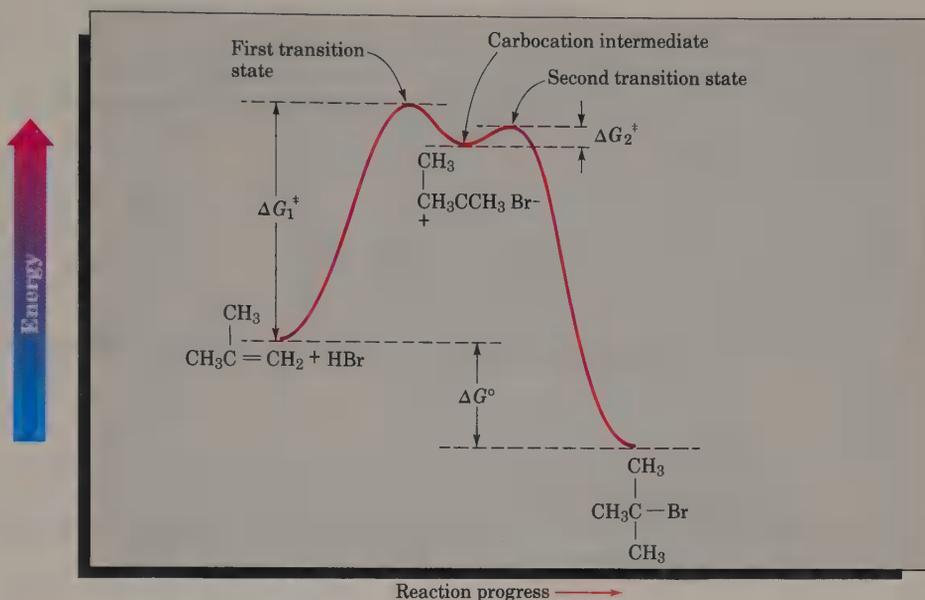
**Figure 6.11** Mechanism of the electrophilic addition of HBr to 2-methylpropene. The reaction occurs in two steps and involves a carbocation intermediate.

shown by the curved arrow at the top of Figure 6.11. The carbocation intermediate that results is itself an electrophile that can accept an electron pair from nucleophilic bromide ion to form a C–Br bond and yield a neutral addition product.

The reaction energy diagram for the overall electrophilic addition reaction, previously shown in Figure 5.6 and repeated in Figure 6.12, has two peaks (transition states) separated by a valley (carbocation intermediate). The energy level of this intermediate is higher than that of the starting alkene, but the reaction as a whole is exothermic (negative  $\Delta G^\circ$ ). The first step, protonation of the alkene to yield the intermediate cation, is relatively slow, but once formed, the cation intermediate rapidly reacts further to yield the final bromoalkane product. The relative rates of the two individual steps are indicated in Figure 6.12 by the fact that  $\Delta G_1^\ddagger$  is larger than  $\Delta G_2^\ddagger$ .

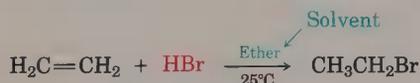
## 6.9 Addition of HX to Alkenes: Carbocations

Electrophilic addition of HX to alkenes is a general reaction that makes it possible to synthesize a variety of products. For example, addition of HCl and HBr is straightforward:<sup>3</sup>

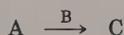


**Figure 6.12** Reaction energy diagram for the two-step electrophilic addition of HBr to 2-methylpropene. The first step is slower than the second step.

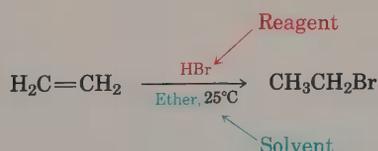
<sup>3</sup>Organic reaction equations can be written in different ways to emphasize different points. For example, the reaction of ethylene with HBr might be written in the format  $A + B \rightarrow C$  to emphasize that both reaction partners are equally important for the purposes of the discussion. The reaction solvent and notes about other reaction conditions such as temperature are usually written either above or below the reaction arrow.



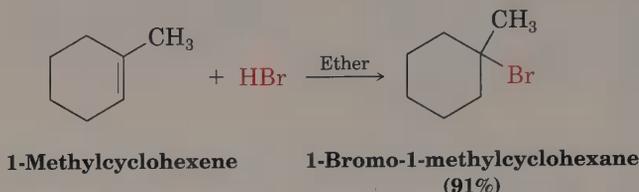
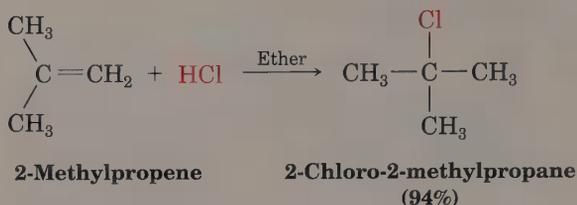
Alternatively, we might choose to write the same reaction in the format



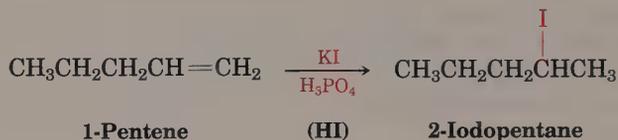
to emphasize that reagent A is the organic starting material whose chemistry is of greater interest. Reagent B is then placed above the reaction arrow, together with notes about solvent and reaction conditions. For example:



Both reaction formats are frequently used in chemistry, and you sometimes have to look carefully at the overall transformation to see what the different roles of the chemicals shown next to the reaction arrows are.



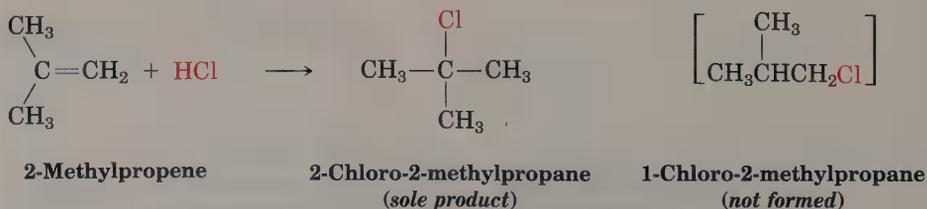
Addition of HI to alkenes also occurs, but it's best to use a mixture of phosphoric acid and potassium iodide to generate HI in the reaction mixture, rather than to use HI directly. The overall mechanism is the same as for the other additions:



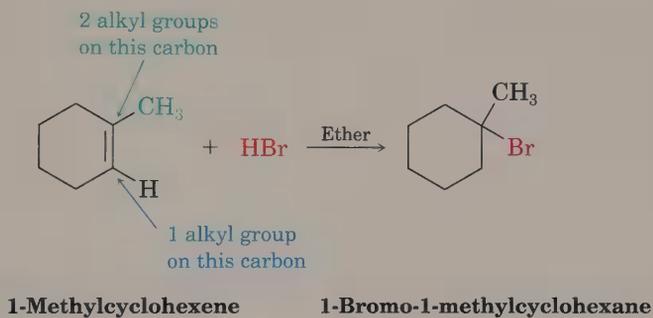
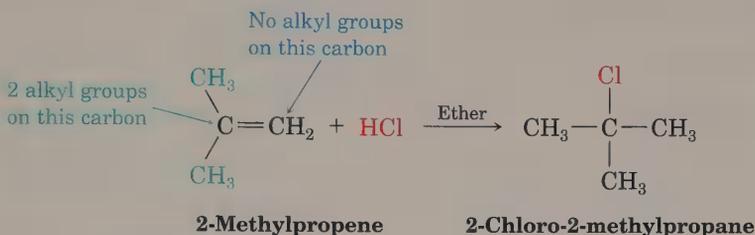
## 6.10 Orientation of Electrophilic Addition: Markovnikov's Rule

Look carefully at the three reactions just shown. In all cases, an unsymmetrically substituted alkene has given a *single* addition product, rather than the mixture that might have been expected. For example, 2-methylpropene might have reacted with HCl to give 1-chloro-2-methylpropane (isobutyl chloride) in addition to 2-chloro-2-methylpropane, but it didn't. We say that such reactions are **regiospecific** (ree-jee-oh-specific) when only one of two possible directions of addition is observed.

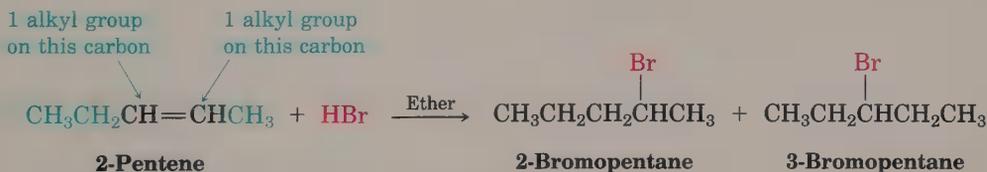
A regiospecific reaction:



By looking at the results of many such reactions and cataloging them, the Russian chemist Vladimir Markovnikov<sup>4</sup> proposed in 1869 what has become known as **Markovnikov's rule**: *In the addition of HX to an alkene, the H attaches to the carbon with fewer alkyl substituents, and the X attaches to the carbon with more alkyl substituents*:

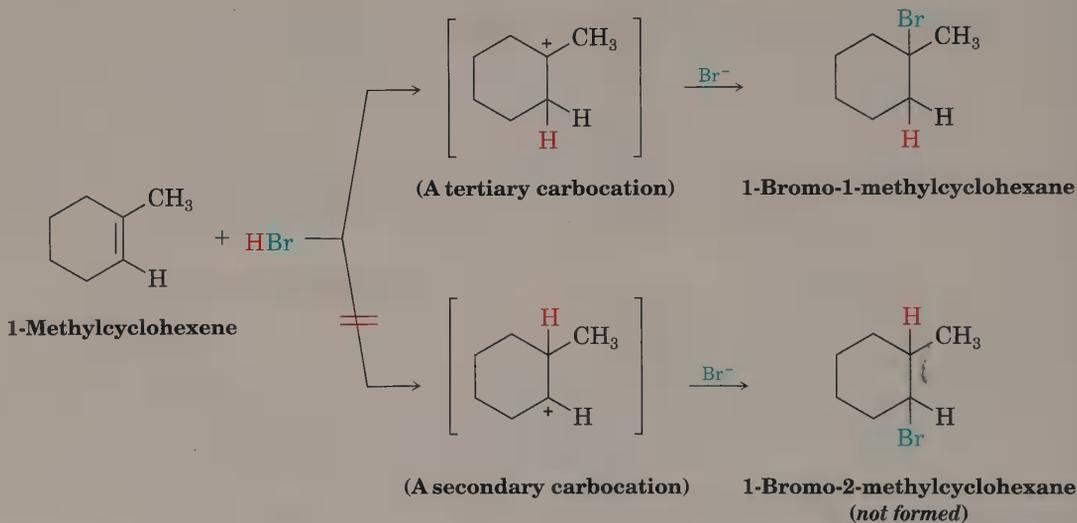
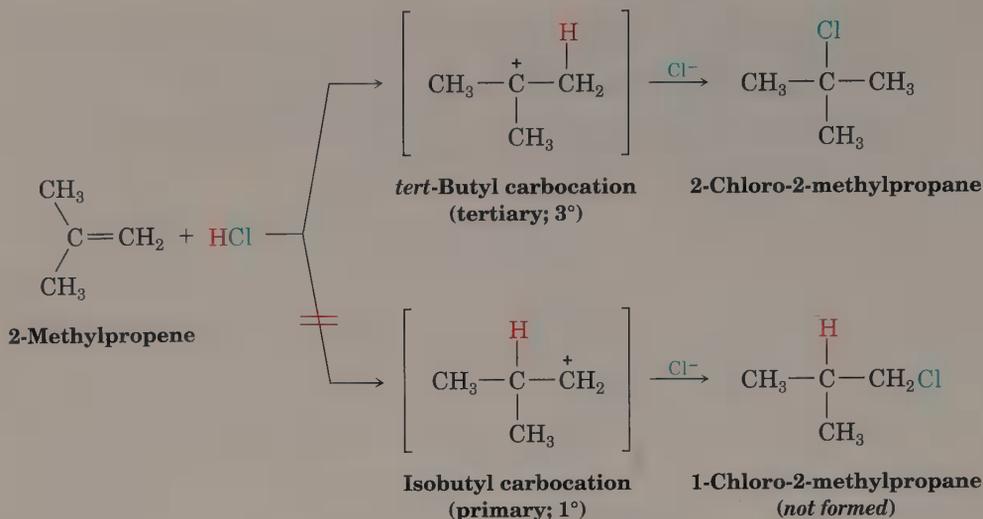


When both ends of the double bond have the same degree of substitution, a mixture of products results:



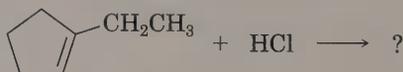
Since carbocations are involved as intermediates in these reactions, another way to express Markovnikov's rule is to say that, in the addition of HX to an alkene, the more highly substituted carbocation is formed as an intermediate in preference to the less highly substituted one. For example, addition of H<sup>+</sup> to 2-methylpropene yields the intermediate *tertiary* carbocation rather than the primary carbocation, and addition to 1-methylcyclohexene yields a tertiary cation rather than a secondary one. Why should this be?

<sup>4</sup>Vladimir Vassilyevich Markovnikov (or Markownikoff) (1833–1904); b. Nijni-Novgorod, Russia; pupil of Butlerov, Erlenmeyer, Baeyer, and Kolbe; professor in Odessa (1871) and Moscow (1873).

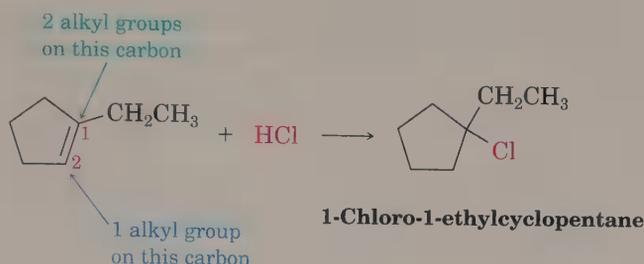


**PRACTICE PROBLEM**.....

What product would you expect from reaction of HCl with 1-ethylcyclopentene?

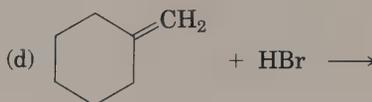
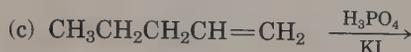
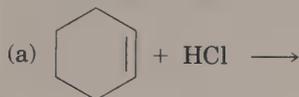


**Solution** Markovnikov's rule predicts that the hydrogen will add to the double-bond carbon that has one alkyl group (C2 on the ring) and the chlorine will add to the double-bond carbon that has two alkyl groups (C1 on the ring). The expected product is 1-chloro-1-ethylcyclopentane.



PROBLEM.....

6.14 Predict the products of these reactions:



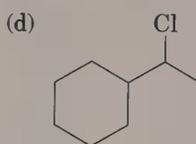
PROBLEM.....

6.15 What alkenes would you start with to prepare these alkyl halides?

(a) Bromocyclopentane

(b)  $\text{CH}_3\text{CH}_2\text{CHBrCH}_2\text{CH}_2\text{CH}_3$

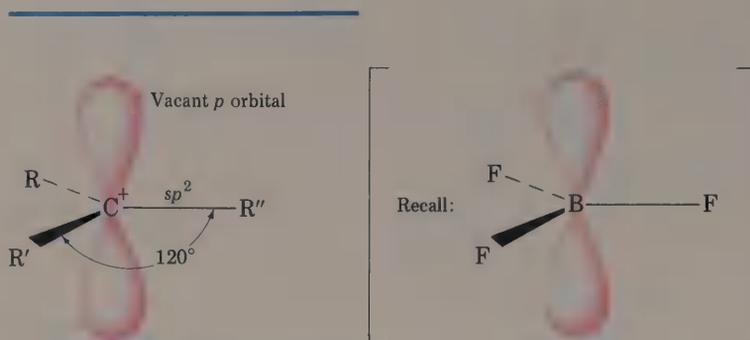
(c) 1-Iodo-1-ethylcyclohexane



## 6.11 Carbocation Structure and Stability

To understand the reasons for the Markovnikov orientation of electrophilic addition reactions, we need to learn more about the structure and stability of carbocations and about the general nature of reactions and transition states. The first point to explore involves structure.

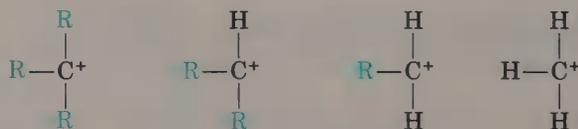
A great deal of evidence has shown that carbocations are *planar*; the carbon is  $sp^2$  hybridized and the three substituents are oriented to the corners of an equilateral triangle, as indicated in Figure 6.13. Since there are only six valence electrons on carbon, and since all six are used in the three sigma bonds, the  $p$  orbital extending above and below the plane is unoccupied. (Note the electronic similarity of carbocations to trivalent boron compounds such as  $\text{BF}_3$  [Section 1.13].)



**Figure 6.13** The electronic structure of a carbocation. The carbon is  $sp^2$  hybridized and has a vacant  $p$  orbital.

The second point to explore involves carbocation stability. 2-Methylpropene might react with HCl to form a carbocation having three alkyl substituents (a tertiary ion,  $3^\circ$ ), or it might react to form a carbocation having one alkyl substituent (a primary ion,  $1^\circ$ ). Since the tertiary chloride, 2-chloro-2-methylpropane, is the only product observed, formation of the tertiary cation is evidently favored over formation of the primary cation. Thermodynamic measurements show that, indeed, the stability of carbocations increases with increasing substitution: More highly substituted carbocations are more stable than less highly substituted ones.

One way of determining carbocation stabilities is to measure the amount of energy required to form the carbocation from its corresponding alkyl halide,  $R-X \rightarrow R^+ + :X^-$ . As shown in Table 6.5 and Figure 6.14, tertiary halides ionize to give carbocations much more readily than secondary or primary halides. As a result, trisubstituted (tertiary,  $3^\circ$ ) carbocations are more stable than disubstituted (secondary,  $2^\circ$ ) ones, which are more stable than monosubstituted (primary,  $1^\circ$ ) ones:



Tertiary ( $3^\circ$ ) > Secondary ( $2^\circ$ ) > Primary ( $1^\circ$ ) > Methyl

More stable



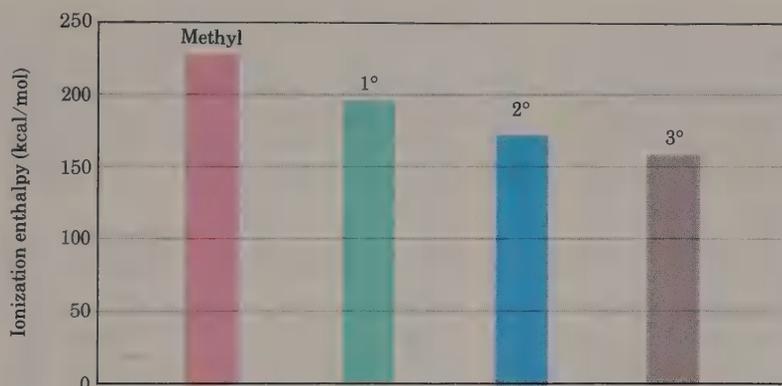
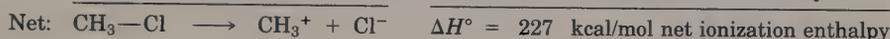
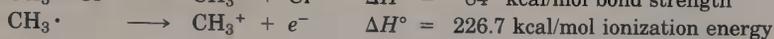
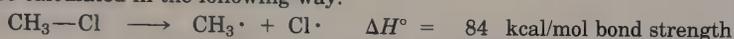
Less stable

Although the data in Table 6.5 are taken from measurements made in the gas phase, a similar carbocation stability order is found in solution. The values for ionization are much lower in solution since polar solvents can stabilize the ions, but the order of carbocation stability remains the same.

TABLE 6.5 Enthalpies of Ionization for Alkyl Halides in the Gas Phase:<sup>a</sup> R-Cl → R<sup>+</sup> + Cl<sup>-</sup>

Type	Reaction	$\Delta H^\circ$	
		(kcal/mol)	(kJ/mol)
Methyl	$\text{CH}_3\text{Cl} \longrightarrow \text{CH}_3^+ + \text{Cl}^-$	227	950
Primary	$\text{CH}_3\text{CH}_2\text{Cl} \longrightarrow \text{CH}_3\text{CH}_2^+ + \text{Cl}^-$	195	816
Secondary	$(\text{CH}_3)_2\text{CHCl} \longrightarrow (\text{CH}_3)_2\text{CH}^+ + \text{Cl}^-$	173	724
Tertiary	$(\text{CH}_3)_3\text{CCl} \longrightarrow (\text{CH}_3)_3\text{C}^+ + \text{Cl}^-$	157	657

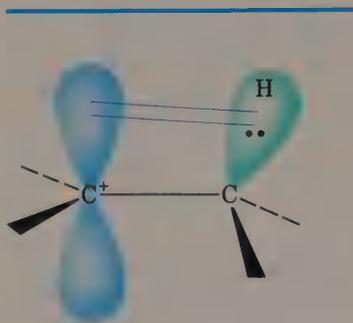
<sup>a</sup>Enthalpies are calculated in the following way:



**Figure 6.14** A graph of ionization enthalpy versus substitution pattern for the gas-phase ionization of alkyl chlorides to yield carbocations. More highly substituted alkyl halides ionize more readily than less highly substituted ones.

Why are more highly substituted carbocations more stable than less highly substituted ones? There are at least two reasons. Part of the answer has to do with hyperconjugation, and part has to do with inductive effects. Hyperconjugation, discussed in Section 6.7 in connection with the stability order of substituted alkenes, is the overlap of a *p* orbital and a neighboring C-H sigma bond orbital. In the present situation, hyperconjugation between

the vacant carbocation  $p$  orbital and a neighboring C–H sigma bond acts to lower the energy level of the carbocation (Figure 6.15).



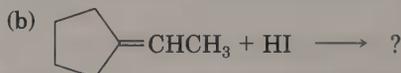
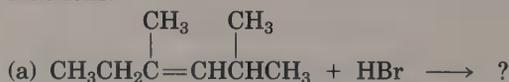
**Figure 6.15** Stabilization of a carbocation through hyperconjugation. The sigma electrons in the neighboring C–H bond help stabilize the positive charge.

The effect of hyperconjugation in carbocations is to allow neighboring sigma bond electrons to stabilize the positive charge by spreading the charge out, or *delocalizing* it, over a greater volume of space. We'll see repeatedly in later chapters that delocalizing charge invariably leads to greater stability. The more alkyl groups there are on the carbocation, the more possibilities there are for hyperconjugation and the more stable the carbocation—in other words: tertiary > secondary > primary > methyl.

Inductive effects, discussed in Section 2.4 in connection with polar covalent bonds, result from the shifting of electrons in a bond in response to the electronegativity of a nearby atom. In the present instance, electrons from a relatively large and polarizable alkyl group can shift toward a neighboring positive charge more easily than electrons from an attached hydrogen can. Thus, the more alkyl groups there are attached to the positively charged carbon, the more electron density shifts toward the charge and the more inductive stabilization of the cation there is.

**PROBLEM** .....

- 6.16** Show the structures of the carbocation intermediates you would expect in these reactions:



.....

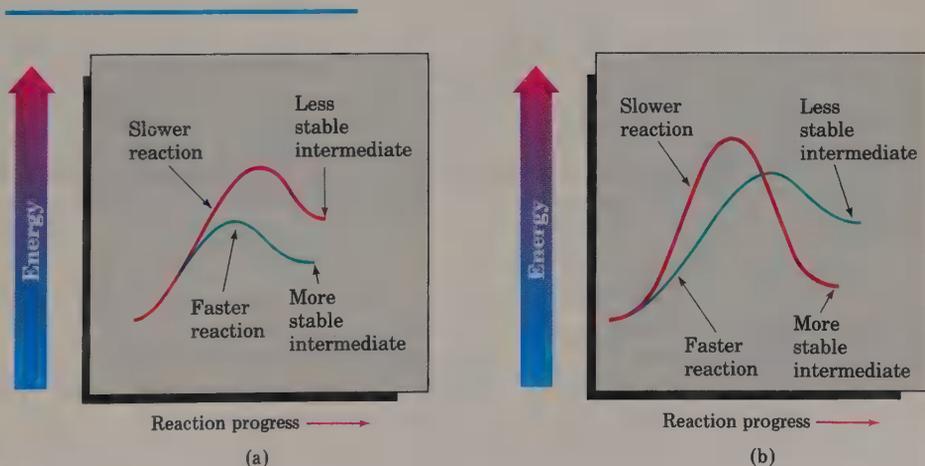
## 6.12 The Hammond Postulate

To summarize our knowledge of electrophilic addition reactions up to this point, we know two facts:

1. We know that electrophilic addition reactions to unsymmetrically substituted alkenes involve the more highly substituted carbocation. A more highly substituted carbocation evidently forms faster than a less highly substituted one and, once formed, rapidly goes on to give the final product.
2. We know that more highly substituted carbocations are more stable than less highly substituted ones. That is, the stability order of carbocations is tertiary > secondary > primary > methyl.

What we haven't yet seen is how these two facts are related. Why is it that the *stability* of the carbocation intermediate affects the *rate* at which the carbocation is formed and thereby determines the structure of the final product? After all, carbocation stability is determined by  $\Delta G^\circ$ , but reaction rate is determined by  $\Delta G^\ddagger$  (activation energy). The two quantities aren't directly related.

Although there is no precise thermodynamic relationship between the stability of a high-energy carbocation intermediate and the rate of its formation, there *is* an intuitive relationship between the two. It's generally true when comparing two similar reactions that the more stable intermediate usually forms faster than the less stable one. The situation is shown graphically in Figure 6.16, where the reaction energy profile shown in part

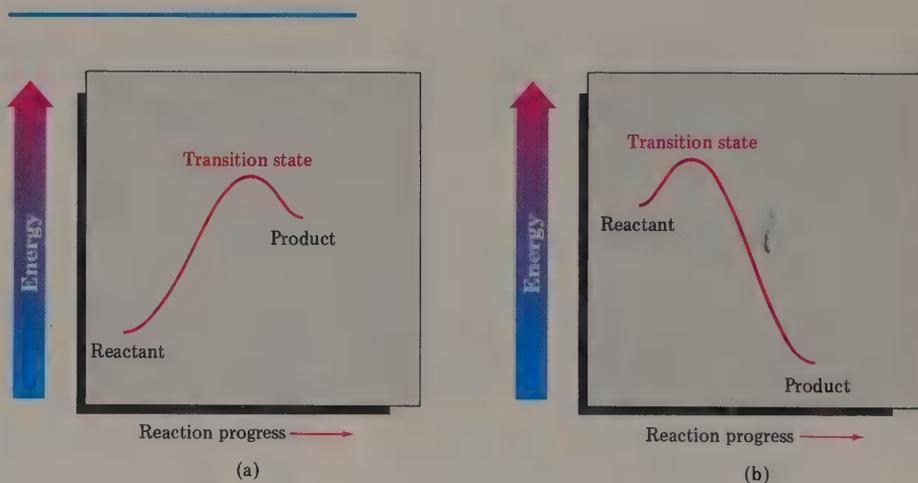


**Figure 6.16** Reaction energy diagrams for two similar competing reactions. In (a), the faster reaction yields the more stable intermediate. In (b), the slower reaction yields the more stable intermediate. The curve shown in (a) represents the usual situation.

(a) represents the usual situation rather than the profile shown in part (b). That is, the two curves remain on parallel courses rather than crossing each other.

An explanation of the relationship between reaction rate and intermediate stability was first advanced in 1955. Known as the **Hammond postulate**,<sup>5</sup> this explanation is not a thermodynamic law; it is simply a reasonable account of observed facts. It intuitively links reaction rate and intermediate stability by looking at the energy level and structure of the transition state.

Transition states represent energy maxima. They are high-energy activated complexes that occur transiently during the course of a reaction and that immediately go on to a more stable species. Although we can't actually *observe* transition states, because they have no finite lifetime, the Hammond postulate says that we can get an *idea* of the structure of a particular transition state by looking at the structure of the stable species in the nearest energy minimum. In terms of reaction energy diagrams, we can imagine the two cases shown in Figure 6.17. The reaction profile in part (a) shows the energy curve for an endothermic reaction step, and the profile in part (b) shows the curve for an exothermic step.



**Figure 6.17** Reaction energy diagrams for endothermic and exothermic steps. (a) In an endothermic step, the energy levels of transition state and *product* are similar. (b) In an exothermic step, the energy levels of transition state and *reactant* are similar.

<sup>5</sup>George Simms Hammond (1921– ); b. Auburn, Maine; Ph.D. (1947), Harvard University; professor, Iowa State University; California Institute of Technology; University of California, Santa Cruz; Allied Chemical Company.

In the endothermic reaction in Figure 6.17(a), the energy level of the transition state is much closer to that of the product than to that of the reactant. Since the transition state is closer *energetically* to the product, we make the natural assumption that it's also closer *structurally*. In other words, we can say that *the transition state for an endothermic reaction step structurally resembles the product of that step*. Conversely, the transition state for the exothermic reaction in part (b) is much closer energetically to the reactant than to the product, and we say that *the transition state for an exothermic reaction step structurally resembles the reactant for that step*.

### Hammond postulate

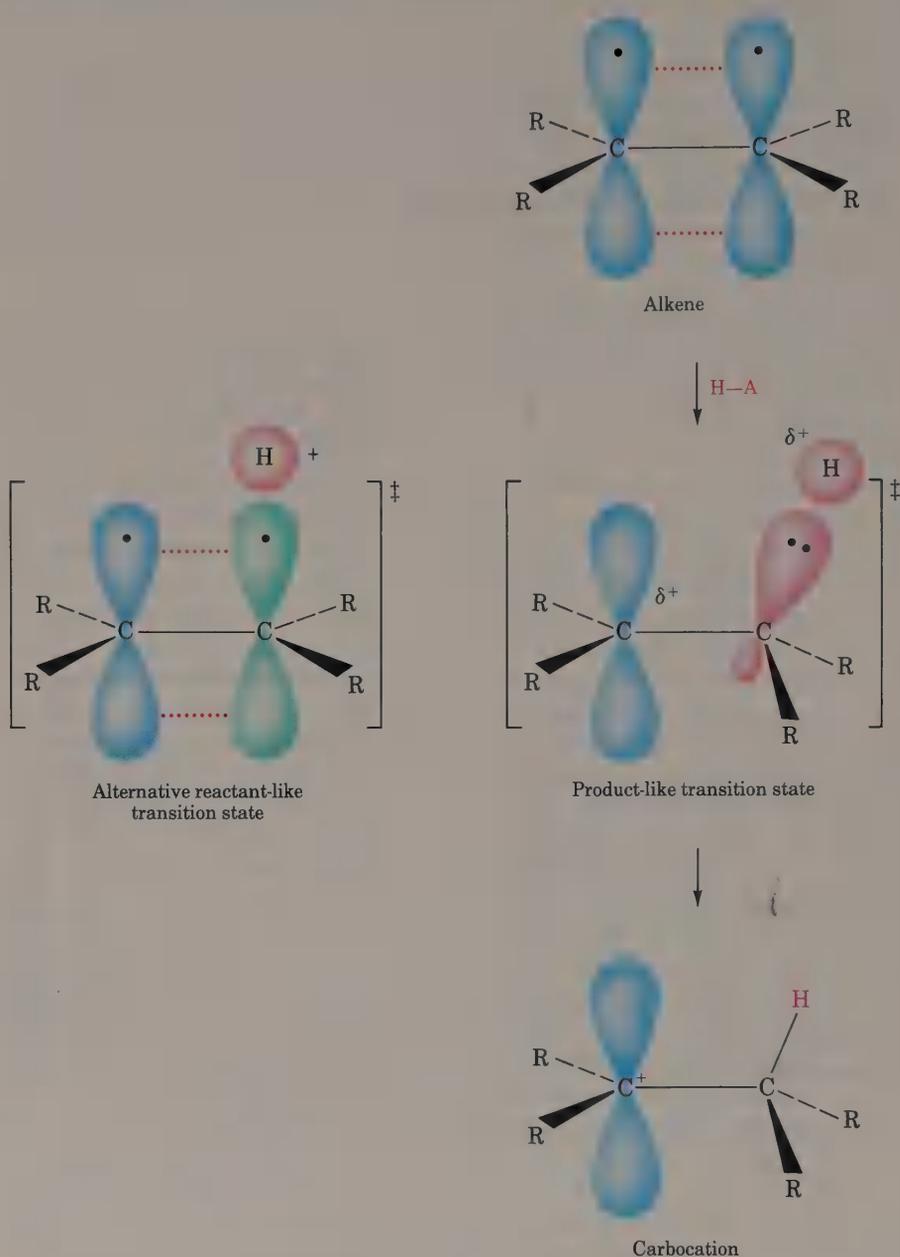
The structure of a transition state resembles the structure of the nearest stable species. Transition states for endothermic steps structurally resemble products, and transition states for exothermic steps structurally resemble starting materials.

How does the Hammond postulate apply to electrophilic additions to alkenes? We know that the formation of a carbocation by protonation of an alkene is an endothermic step. Therefore, the transition state for alkene protonation should structurally resemble the carbocation intermediate, and any factor that makes the carbocation product more stable should also make the nearby transition state more stable. Since increasing alkyl substitution stabilizes carbocations, it also stabilizes the transition states leading to those ions, thus resulting in faster reaction. More highly substituted carbocations form faster because their stability is reflected in the transition state that forms them. A hypothetical transition state for alkene protonation might be expected to look like that shown in Figure 6.18.

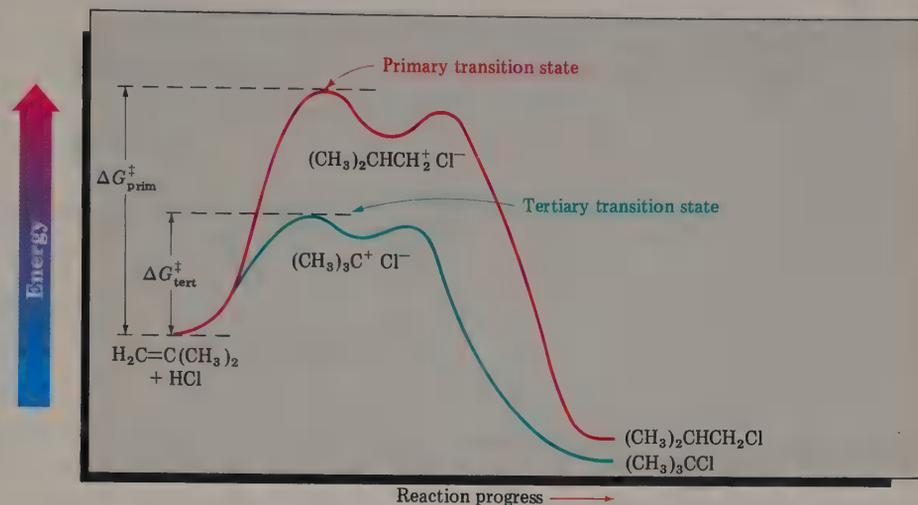
Since the transition state for alkene protonation shown in Figure 6.18 resembles the carbocation product, we can imagine it to be a structure in which one of the alkene carbon atoms has almost completely rehybridized from  $sp^2$  to  $sp^3$  and in which the remaining alkene carbon bears a substantial portion of the positive charge. The positive charge in this transition state is delocalized and stabilized by hyperconjugation in the same way that the product carbocation is. The more alkyl groups that are present, the greater the extent of charge stabilization in the transition state and the faster it forms. Figure 6.19 summarizes the situation by showing competing reaction energy profiles for the reaction of 2-methylpropene with HCl.

#### PROBLEM .....

- 6.17** What about the second step in the electrophilic addition of HCl to an alkene—the reaction of chloride ion with the carbocation intermediate? Is this step exothermic or endothermic? According to the Hammond postulate, should the transition state for this second step resemble the reactant (carbocation) or product (chloroalkane)? Make a rough drawing of what you would expect the transition-state structure to look like.
- .....



**Figure 6.18** The structure of a hypothetical transition state for alkene protonation. The transition state is closer in both energy and structure to the carbocation than to the reactant. Thus, an increase in carbocation stability (lower  $\Delta G^\circ$ ) also causes an increase in transition-state stability (lower  $\Delta G^\ddagger$ ).



**Figure 6.19** A reaction energy diagram for the electrophilic addition of HCl to 2-methylpropene. The tertiary cation forms faster than the primary cation intermediate because it's more stable. The same factors that make the tertiary cation more stable also make the transition state leading to it more stable.

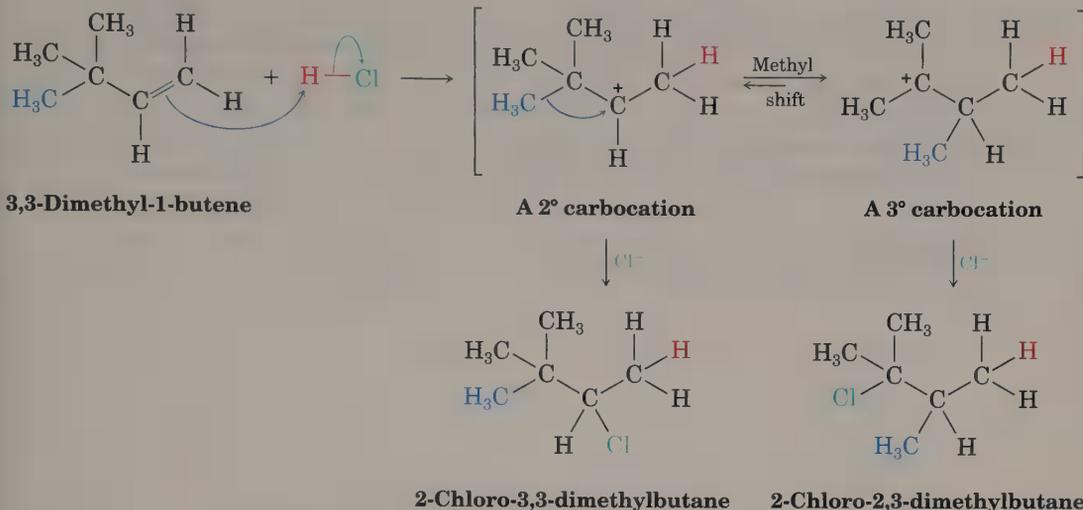
## 6.13 Mechanistic Evidence: Carbocation Rearrangements

How do we know that the carbocation mechanism for addition of HX to alkenes is correct? The answer is that we *don't* know it's correct, or at least we don't know with absolute certainty. Although an *incorrect* reaction mechanism can be *disproved* by demonstrating that it doesn't satisfactorily account for observed data, a correct reaction mechanism can never be entirely proven. The best we can do is to show that a proposed mechanism is fully consistent with all known facts. If enough data are satisfactorily explained, then the mechanism is probably correct.

What evidence is there to support the two-step carbocation mechanism we've proposed for the reaction of HX with alkenes? How do we know that the two reactants, HX and alkene, don't simply come together in a single step to give the final product without going through a carbocation intermediate? One of the best pieces of evidence for a carbocation mechanism was discovered during the 1930s by F. C. Whitmore<sup>6</sup> at Pennsylvania State

<sup>6</sup>Frank C. Whitmore (1887–1947); b. North Attleboro, Mass.; Ph.D. Harvard (E. L. Jackson); professor, Pennsylvania State University.

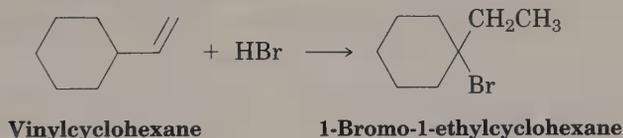




Note the similarities of these two carbocation rearrangements. In both cases, a group ( $\text{:H}^-$  or  $\text{:CH}_3^-$ ) moves to a positively charged carbon, *taking its electron pair with it*. Also in both cases, a less stable carbocation rearranges to a more stable ion. Rearrangements of the sort just shown are a common feature of carbocation chemistry. We'll see at numerous places in future chapters that their occurrence in a reaction provides strong mechanistic evidence for the presence of carbocation intermediates.

PROBLEM .....

- 6.18 Propose a mechanism to account for the formation of 1-bromo-1-ethylcyclohexane on reaction of vinylcyclohexane with HBr.

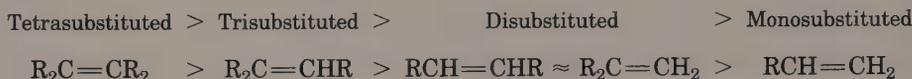


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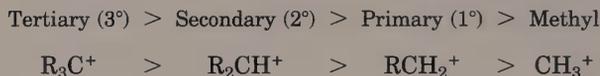
## 6.14 Summary and Key Words

**Alkenes** are hydrocarbons that contain one or more carbon-carbon double bonds. Because they contain fewer hydrogens than related alkanes, alkenes are often referred to as being **unsaturated**. A double bond has two parts: a sigma bond formed by head-on overlap of two  $sp^2$  orbitals and a pi bond formed by sideways overlap of two  $p$  orbitals. The overall bond strength of an alkene double bond is greater than that of a carbon-carbon single bond, with the pi bond part estimated at 64 kcal/mol (268 kJ/mol).

Alkenes are named by IUPAC rules using the suffix *-ene*. Because rotation around the double bond is restricted, substituted alkenes can exist as a pair of **cis-trans stereoisomers**. The geometry of a double bond can be specified by application of the Cahn–Ingold–Prelog **sequence rules**, which assign priorities to double-bond substituents. If the high-priority groups on each carbon are on the same side of the double bond, the geometry is *Z* (*zusammen*, “together”); if the high-priority groups on each carbon are on opposite sides of the double bond, the geometry is *E* (*entgegen*, “apart”). The stability order of alkyl-substituted double bonds is:



The chemistry of alkenes is dominated by **electrophilic addition reactions**. When HX reacts with an unsymmetrically substituted alkene, **Markovnikov's rule** predicts that the hydrogen will add to the carbon having fewer alkyl substituents and that the X group will add to the carbon having more alkyl substituents. Electrophilic additions to alkenes take place through **carbocation** intermediates formed by reaction of the nucleophilic alkene pi bond with electrophilic H–A. Carbocation stability follows the order



Markovnikov's rule can be restated by saying that, in the addition of HX to an alkene, the more stable carbocation intermediate is formed. This result is explained by the **Hammond postulate**, which states that the transition state of an exothermic reaction step structurally resembles the reactant, whereas the transition state of an endothermic reaction step structurally resembles the product. Since an alkene protonation step is endothermic, the stability of the more highly substituted carbocation is reflected in the stability of the transition state leading to its formation.

One of the best pieces of evidence in support of a carbocation mechanism for electrophilic addition reactions is the observation of structural **rearrangements** that sometimes occur during reaction. Rearrangements occur by shift of either a hydride ion,  $\text{:H}^-$  (a **hydride shift**), or an alkyl group anion,  $\text{:R}^-$ , from a neighboring carbon atom to the positively charged carbon. The net result is isomerization of a less stable carbocation to a more stable one.

## ADDITIONAL PROBLEMS .....

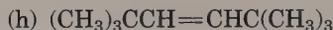
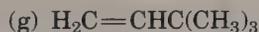
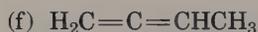
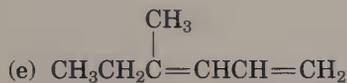
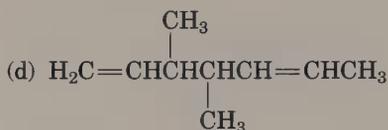
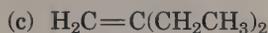
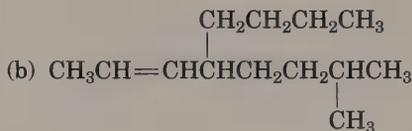
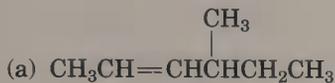
- 6.19 Calculate the number of double bonds and/or rings in these formulas:
- |   |   |
|---|---|
| (a) Benzene, $\text{C}_6\text{H}_6$                 | (b) Cyclohexene, $\text{C}_6\text{H}_{10}$          |
| (c) Myrcene (bay oil), $\text{C}_{10}\text{H}_{16}$ | (d) Lindane, $\text{C}_6\text{H}_6\text{Cl}_6$      |
| (e) Pyridine, $\text{C}_5\text{H}_5\text{N}$        | (f) Safrole, $\text{C}_{10}\text{H}_{10}\text{O}_2$ |

6.20 Calculate the number of multiple bonds and/or rings in these formulas, and then draw five possible structures for each:

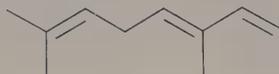
- (a)  $C_{10}H_{16}$  (b)  $C_8H_8O$  (c)  $C_7H_{10}Cl_2$   
 (d)  $C_{10}H_{16}O_2$  (e)  $C_5H_9NO_2$  (f)  $C_3H_10ClNO$

6.21 A compound of formula  $C_{10}H_{14}$  undergoes catalytic hydrogenation but absorbs only two equivalents of hydrogen. How many rings does the compound have?

6.22 Give IUPAC names for these alkenes:



6.23 Ocimene is a triene found in the essential oil of many plants. What is its correct IUPAC name?



Ocimene

6.24  $\alpha$ -Farnesene is a constituent of the natural wax found on apples. What is its IUPAC name?



$\alpha$ -Farnesene

6.25 Indicate *E* or *Z* stereochemistry for each of the double bonds in  $\alpha$ -farnesene (Problem 6.24).

6.26 Draw structures corresponding to these systematic names:

- (a) 2,4-Dimethyl-1,4-hexadiene (b) 3,3-Dimethyl-4-propyl-1,5-octadiene  
 (c) 4-Methyl-1,2-pentadiene (d) 2,6-Dimethyl-1,3,5,7-octatetraene  
 (e) 3-Butyl-2-heptene (f) 2,2,5,5-Tetramethyl-3-hexene

6.27 Menthene, a hydrocarbon found in mint plants, has the systematic name 1-isopropyl-4-methylcyclohexene. Draw its structure.

6.28 These names are incorrect. Draw structures and give correct IUPAC names:

- (a) 2-Methyl-2,4-pentadiene (b) 3-Methylene-1-pentene  
 (c) 3,6-Octadiene (d) 5-Ethyl-4-octene  
 (e) 3-Propyl-3-heptene (f) 3-Vinyl-1-propene

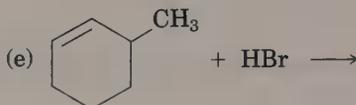
6.29 Draw and name the five possible pentene isomers,  $C_5H_{10}$ . Ignore *E,Z* isomers.

- 6.30 Draw and name the 13 possible hexene isomers,  $C_6H_{12}$ . Ignore *E,Z* isomers.
- 6.31 Which of the compounds you drew in Problems 6.29 and 6.30 show *cis-trans* isomerism?
- 6.32 According to heat-of-hydrogenation data, *trans*-2-butene is more stable than *cis*-2-butene by 1 kcal/mol. Hydrogenation measurements also show that *trans*-2,2,5,5-tetramethyl-3-hexene is more stable than its *cis* isomer by 9.3 kcal/mol. Explain this large difference.

		$\Delta H^\circ_{\text{hydrog}}$	
		(kcal/mol)	(kJ/mol)
Cis	$CH_3CH=CHCH_3$	28.6	120
Trans	$CH_3CH=CHCH_3$	27.6	116
Cis	$(CH_3)_3CCH=CHC(CH_3)_3$	36.2	151
Trans	$(CH_3)_3CCH=CHC(CH_3)_3$	26.9	112

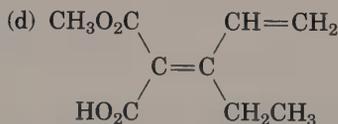
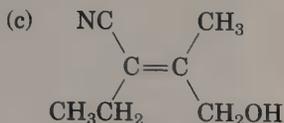
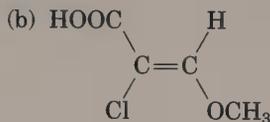
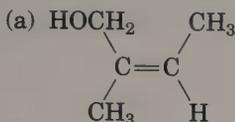
- 6.33 Allene (1,2-propadiene),  $H_2C=C=CH_2$ , has two adjacent double bonds. What kind of hybridization must the central carbon have? Sketch the bonding orbitals in allene. What shape do you predict for allene?
- 6.34 1,4-Pentadiene, a compound with two nonadjacent double bonds, has  $\Delta H^\circ_{\text{hydrog}} = 60.8$  kcal/mol, which is approximately twice the value for 1-pentene ( $\Delta H^\circ_{\text{hydrog}} = 30.1$  kcal/mol). 1,2-Pentadiene, however, has  $\Delta H^\circ_{\text{hydrog}} = 71.3$  kcal/mol. What does this tell you about the stability of 1,2-pentadiene? What explanation might there be for these results?
- 6.35 Predict the major product in each of these reactions:

- (a)  $CH_3CH_2CH=C(CH_3)CH_2CH_3 + HCl \longrightarrow$
- (b) 1-Ethylcyclopentene + HBr  $\longrightarrow$
- (c) 2,2,4-Trimethyl-3-hexene + HI  $\longrightarrow$
- (d) 1,6-Heptadiene + 2 HCl  $\longrightarrow$

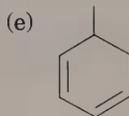
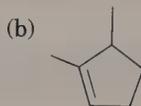
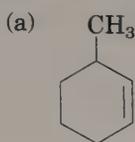


- 6.36 Rank the following sets of substituents in order of priority according to the Cahn-Ingold-Prelog sequence rules:
- (a)  $-CH_3$ ,  $-Br$ ,  $-H$ ,  $-I$
- (b)  $-OH$ ,  $-OCH_3$ ,  $-H$ ,  $-COOH$
- (c)  $-COOH$ ,  $-COOCH_3$ ,  $-CH_2OH$ ,  $-CH_3$
- (d)  $-CH_3$ ,  $-CH_2CH_3$ ,  $-CH_2CH_2OH$ ,  $-C(=O)CH_3$
- (e)  $-CH=CH_2$ ,  $-CN$ ,  $-CH_2NH_2$ ,  $-CH_2Br$
- (f)  $-CH=CH_2$ ,  $-CH_2CH_3$ ,  $-CH_2OCH_3$ ,  $-CH_2OH$

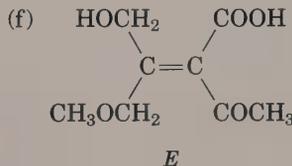
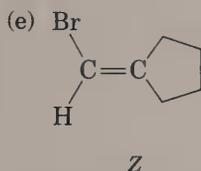
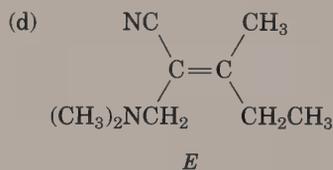
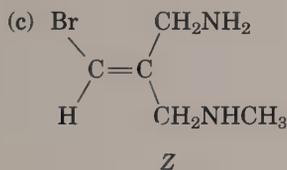
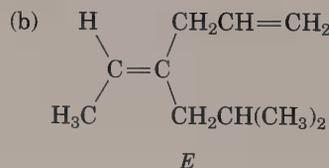
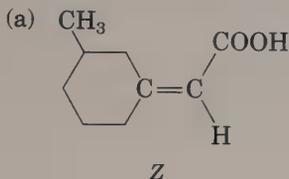
6.37 Assign *E* or *Z* configuration to the following alkenes:



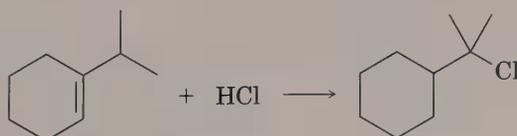
6.38 Name these cycloalkenes according to IUPAC rules:



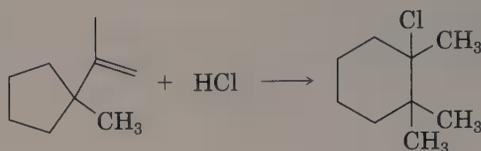
6.39 Which of the given *E*, *Z* designations are correct, and which are incorrect?



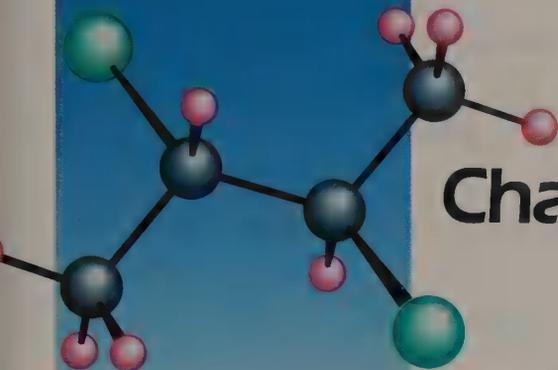
6.40 Propose a mechanism to account for the following reaction. Show the structure of the intermediate(s), and use curved arrows to follow electron flows in each step.



- 6.41 Propose a mechanism to account for the following reaction. Show the structure of the intermediate(s), and use curved arrows to follow electron flows in each step.



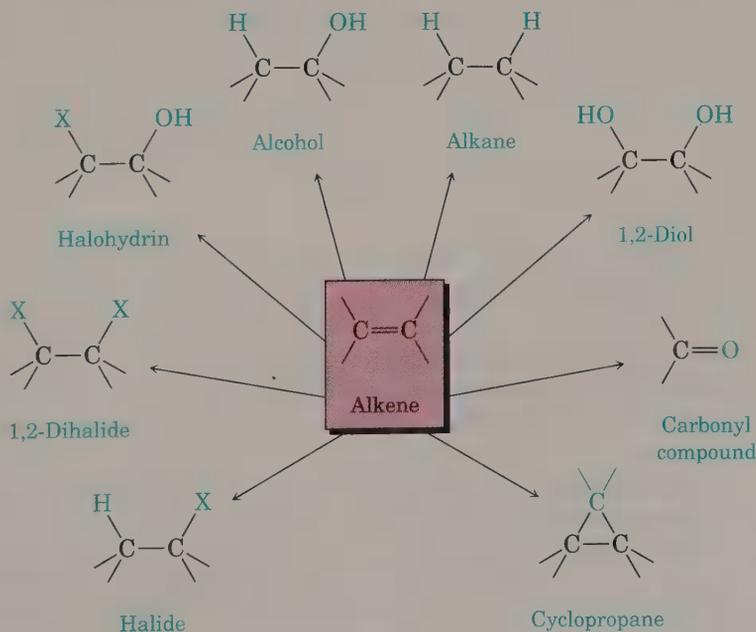
- 6.42 Calculate the degree of unsaturation of these formulas:
- |   |                                  |
|---|----------------------------------|
| (a) Cholesterol, $C_{27}H_{46}O$            | (b) DDT, $C_{14}H_9Cl_5$         |
| (c) Prostaglandin $E_1$ , $C_{20}H_{34}O_5$ | (d) Caffeine, $C_8H_{10}N_4O_2$  |
| (e) Cortisone, $C_{21}H_{28}O_5$            | (f) Atropine, $C_{17}H_{23}NO_3$ |
- 6.43 Draw a reaction energy diagram for the addition of HBr to 1-pentene. Let one curve on your diagram show the formation of 1-bromopentane product and another curve on the same diagram show the formation of 2-bromopentane product. Label the positions for all reactants, intermediates, and products. Which curve has the higher-energy carbocation intermediate? Which curve has the higher first transition state?
- 6.44 Make sketches of the transition-state structures involved in the reaction of HBr with 1-pentene (Problem 6.43). Identify each structure as resembling either starting material or product.



## Chapter 7

# Alkenes: Reactions and Synthesis

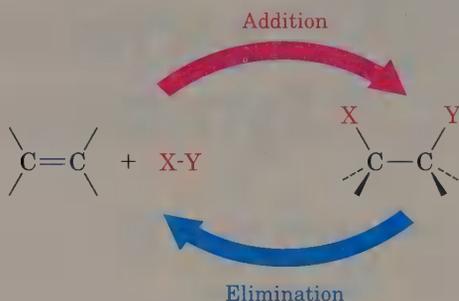
We saw in the preceding chapter that the addition of electrophiles is one of the most important reactions of alkenes. Although we've studied only the addition of HX thus far, many other electrophilic reagents also add to alkenes. In the present chapter, we'll learn how alkenes are prepared, we'll discuss many further examples of alkene addition reactions, and we'll see the wide variety of compound types that can be made from alkenes.



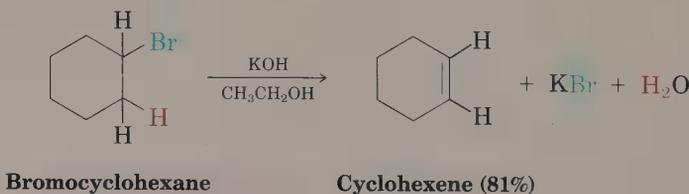
## 7.1 Preparation of Alkenes: A Preview of Elimination Reactions

Before getting to the main topic of this chapter—the reactions of alkenes—let's take a brief look of how alkenes are prepared. We'll return in Chapter 11 for a more detailed look.

Just as the chemistry of alkenes is dominated by addition reactions, the preparation of alkenes is dominated by **elimination reactions**. Additions and eliminations are in many respects two sides of the same coin. That is, an addition reaction might involve the *addition* of HBr or H<sub>2</sub>O to an alkene to form an alkyl halide or an alcohol, whereas an elimination reaction might involve the *loss* of HBr or H<sub>2</sub>O from an alkyl halide or alcohol to form an alkene.

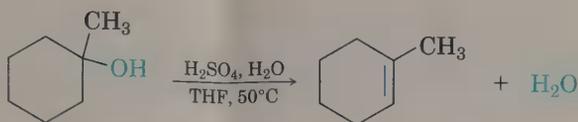


The two most common alkene-forming elimination reactions are **dehydrohalogenation**—the loss of HX from an alkyl halide—and **dehydration**—the loss of water from an alcohol. Dehydrohalogenation usually occurs by reaction of an alkyl halide with strong base such as KOH. For example, bromocyclohexane yields cyclohexene when treated with potassium hydroxide in ethanol solution:



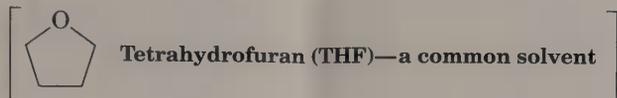
Dehydration is often carried out by treatment of an alcohol with a strong acid. For example, loss of water occurs and 1-methylcyclohexene is formed when 1-methylcyclohexanol is warmed with aqueous sulfuric acid in tetrahydrofuran (THF) solvent.

Both kinds of elimination reactions are sufficiently complex that it's best to defer a detailed discussion until Chapter 11. For the present, it's sufficient just to realize that alkenes are readily available from simple precursors.



1-Methylcyclohexanol

1-Methylcyclohexene (91%)

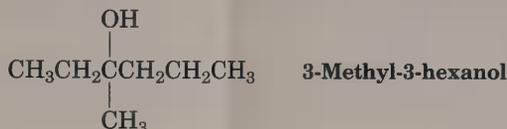


PROBLEM.....

- 7.1 One of the problems in using elimination reactions is that mixtures of products are often formed. For example, treatment of 2-bromo-2-methylbutane with KOH in ethanol yields a mixture of two alkene products. What are their structures?

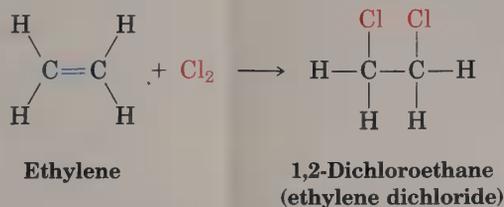
PROBLEM.....

- 7.2 How many alkene products, including *E,Z* isomers, might be obtained by dehydration of 3-methyl-3-hexanol with aqueous sulfuric acid?



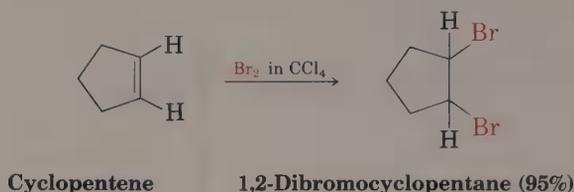
## 7.2 Addition of Halogens to Alkenes

Bromine and chlorine both add readily to alkenes to yield 1,2-dihaloalkanes. For example, more than 7 million tons per year of 1,2-dichloroethane (ethylene dichloride) are synthesized industrially by the addition of  $\text{Cl}_2$  to ethylene. The product is used both as a solvent and as starting material for use in the manufacture of poly(vinyl chloride), PVC.



The addition of bromine to an alkene also serves as a simple and rapid laboratory test for unsaturation. A sample of unknown structure is dissolved

in tetrachloromethane (carbon tetrachloride,  $\text{CCl}_4$ ) and placed in a test tube to which several drops of bromine in  $\text{CCl}_4$  are added. Immediate disappearance of the reddish  $\text{Br}_2$  color signals a positive test and indicates that the sample is an alkene.

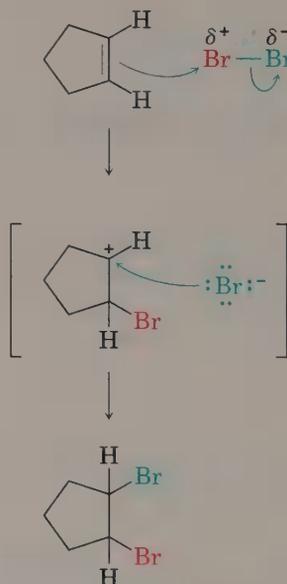


Fluorine tends to be too reactive and difficult to control for most laboratory applications, and iodine does not react with most alkenes.

Based on what we've seen thus far, a possible mechanism for the reaction of halogens with alkenes is shown in Figure 7.1. As a bromine molecule approaches a nucleophilic alkene, the  $\text{Br}-\text{Br}$  bond becomes polarized (Section 5.5). The pi electron pair of the alkene then attacks the positive end of the polarized bromine molecule, displacing bromide ion. The net result is that electrophilic  $\text{Br}^+$  adds to the alkene in the same way that  $\text{H}^+$  adds. The intermediate electrophilic carbocation then immediately reacts with nucleophilic bromide ion to yield the dibromo addition product.

The electron pair from the double bond attacks the polarized bromine, forming a  $\text{C}-\text{Br}$  bond and causing the  $\text{Br}-\text{Br}$  bond to break. Bromide ion departs with both electrons from the former  $\text{Br}-\text{Br}$  bond.

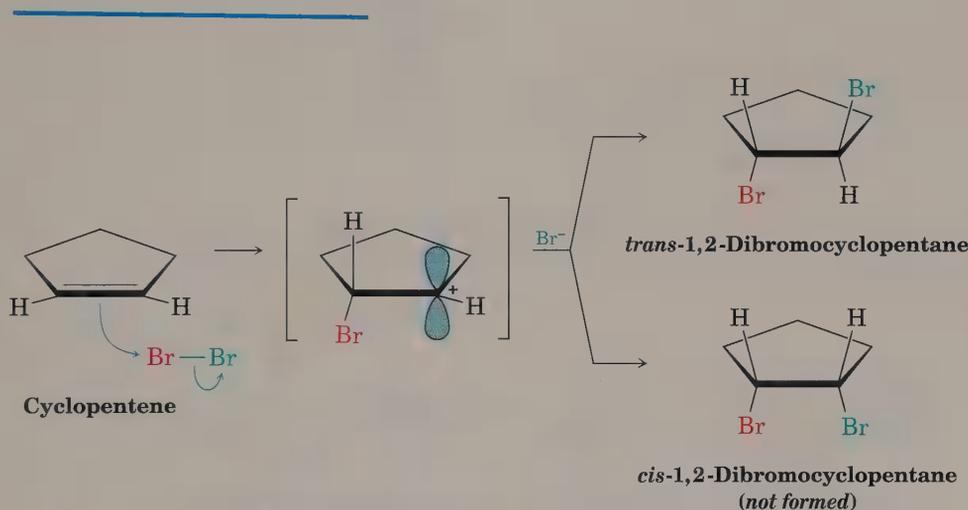
Bromide ion then uses an electron pair to attack the carbocation intermediate, forming a  $\text{C}-\text{Br}$  bond and giving the neutral addition product.



**Figure 7.1** A possible mechanism for the electrophilic addition of  $\text{Br}_2$  to an alkene.

Although the mechanistic description shown in Figure 7.1 for the addition of bromine to alkenes looks reasonable, closer examination shows that it's not completely consistent with known facts. In particular, the proposed mechanism doesn't explain the *stereochemistry* of halogen addition. That is, the mechanism doesn't explain what product stereoisomers (Section 3.9) are formed in the reaction.

Let's look again at the reaction of  $\text{Br}_2$  with cyclopentene and assume that  $\text{Br}^+$  adds to cyclopentene from the bottom face to form the carbocation intermediate shown in Figure 7.2. (The addition could equally well occur from the top face, but we'll consider only one possibility for simplicity.) Since the positively charged carbon is planar and  $sp^2$  hybridized, it could be attacked by bromide ion in the second step of the reaction from either the top or the bottom to give a *mixture* of products. One product has the two bromine atoms on the same side of the ring (cis), whereas the other has the bromines on opposite sides (trans). We find, however, that only *trans*-1,2-dibromocyclopentane is produced. None of the cis product is formed.



**Figure 7.2** Stereochemistry of the addition reaction of bromine with cyclopentene. Only the *trans* product is formed.

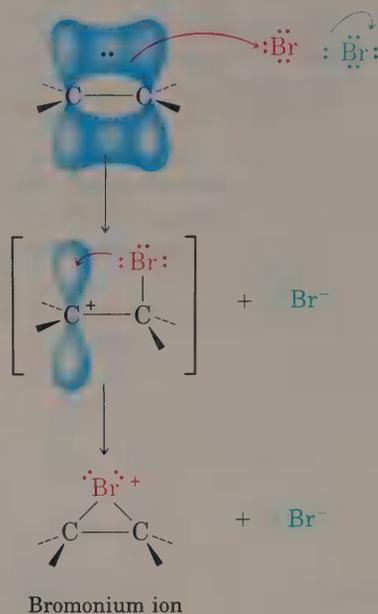
Since the two bromine atoms add to opposite faces of the cyclopentene double bond, we say that the reaction occurs with **anti stereochemistry**. If the two bromines had added from the same face, the reaction would have had **syn stereochemistry**. Note that the word *anti* has a similar meaning in the present stereochemical context to the meaning it has in a butane conformational context (Section 4.3). In both cases, the two groups referred to are  $180^\circ$  apart.

An explanation of anti addition was put forward in 1937 by George Kimball and Irving Roberts, who suggested that the true reaction intermediate is not a carbocation but is instead a **bromonium ion**,  $\text{R}_2\text{Br}^+$ . (A *chloronium ion*, similarly, contains a positively charged, divalent chlorine,  $\text{R}_2\text{Cl}^+$ .) In the present case, the bromonium ion is in a three-membered ring

and is formed by the overlap of bromine lone-pair electrons with the vacant  $p$  orbital of the neighboring carbocation (Figure 7.3). Although Figure 7.3 depicts bromonium ion formation as stepwise, this is done only for clarity. In fact, the bromonium ion is formed in a single step by interaction of the alkene double-bond electrons with  $\text{Br}^+$ .

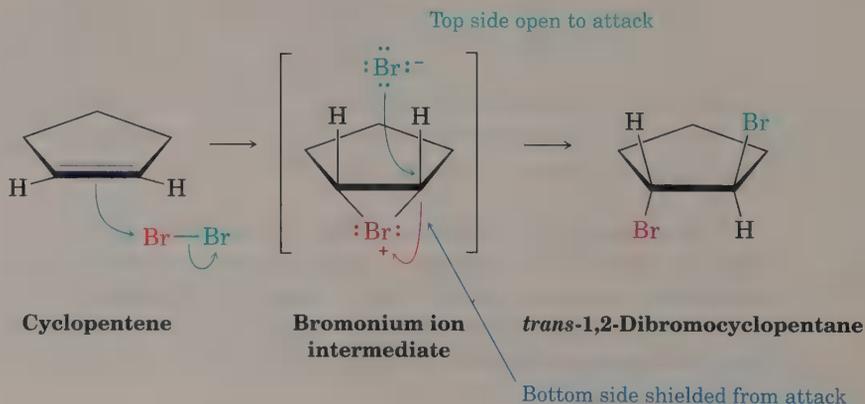
Alkene pi electrons attack bromine, pushing out bromide ion and leaving a bromo carbocation.

The neighboring bromo substituent stabilizes the positive charge by using two of its electrons to overlap the vacant carbon  $p$  orbital, giving a three-membered-ring bromonium ion.

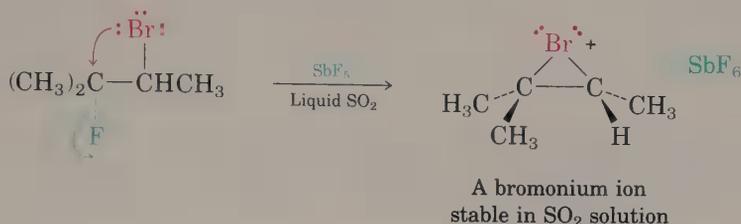


**Figure 7.3** Formation of a bromonium ion intermediate by electrophilic addition of  $\text{Br}^+$  to an alkene.

How does bromonium ion formation account for the anti stereochemistry of addition to cyclopentene? If a bromonium ion is formed as an intermediate, we can imagine that the bromine atom might “shield” one face of the molecule. Attack by bromide ion in the second step could then occur only from the opposite, unshielded face to give anti product.



The halonium ion postulate, made more than 50 years ago to explain the stereochemistry of halogen addition to alkenes, is a remarkable example of deductive logic in chemistry. Arguing from known experimental results, chemists were able to make a hypothesis about the intimate mechanistic details of alkene electrophilic reactions. Much more recently, strong evidence supporting the postulate has come from the work of George Olah, who has observed and studied specially prepared *stable* solutions of cyclic bromonium ions. There's no question that bromonium ions are indeed real.



PROBLEM .....

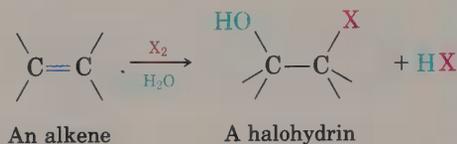
- 7.3 What product would you expect to obtain from addition of  $\text{Cl}_2$  to 1,2-dimethylcyclohexene? Show the stereochemistry of the product.

PROBLEM .....

- 7.4 Unlike the reaction in Problem 7.3, addition of  $\text{HCl}$  to 1,2-dimethylcyclohexene yields a mixture of two products. Show the stereochemistry of each, and explain why a mixture is formed.
- .....

## 7.3 Halohydrin Formation

A great many different kinds of electrophilic additions to alkenes can take place. For example, alkenes can add  $\text{HO}-\text{Cl}$  or  $\text{HO}-\text{Br}$  under suitable conditions to yield 1,2-halo alcohols, or **halohydrins**. Halohydrin formation doesn't take place by direct reaction of an alkene with the reagents  $\text{HOBr}$  or  $\text{HOCl}$ . Rather, the addition is done indirectly by reaction of the alkene with either  $\text{Br}_2$  or  $\text{Cl}_2$  in the presence of water.



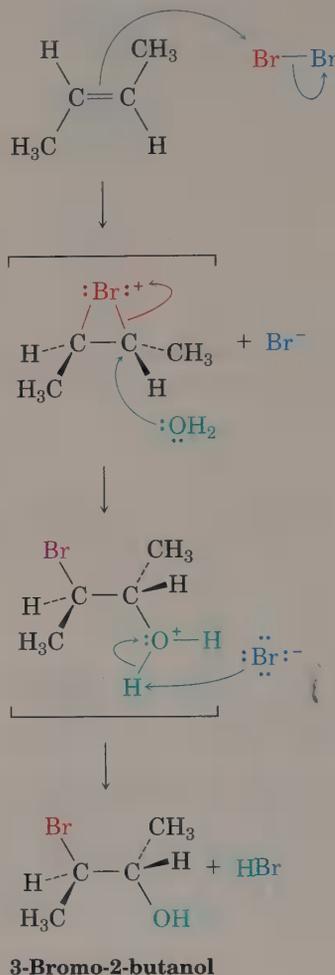
We've seen that, when a solution of bromine in carbon tetrachloride reacts with an alkene, the cyclic bromonium ion intermediate is trapped by the only nucleophile present, bromide ion. If, however, the reaction is carried out *in the presence of an additional nucleophile*, the intermediate bromonium ion can be "intercepted" by the added nucleophile and diverted to a different product. When an alkene reacts with bromine in the presence of water, for

example, water competes with bromide ion as nucleophile and reacts with the bromonium ion intermediate to yield a mixture of dibromide and bromo alcohol (a **bromohydrin**). The net effect is addition of HO–Br to the alkene. The reaction takes place by the pathway shown in Figure 7.4.

Reaction of the alkene with Br<sub>2</sub> yields a bromonium ion intermediate.

Water acts as a nucleophile, using a lone pair of electrons to open the bromonium ion ring and form a bond to carbon. Since oxygen donates its electrons in this step, it now has the positive charge.

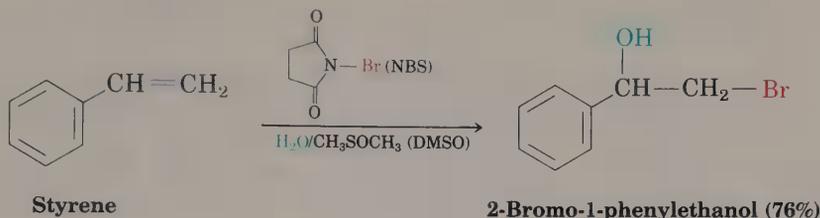
Loss of a proton (H<sup>+</sup>) from oxygen then gives HBr and the neutral bromohydrin addition product.



**Figure 7.4** Mechanism of bromohydrin formation by reaction of an alkene with bromine in the presence of water.

In practice, few alkenes are soluble in water, and bromohydrin formation is often carried out in a solvent such as aqueous dimethyl sulfoxide (DMSO), using a reagent called *N*-bromosuccinimide (NBS) as a bromine source. (NBS is a stable, easily handled compound that slowly decomposes in water to yield Br<sub>2</sub> at a controlled rate. Bromine itself can also be used

in the addition reaction, but it is much more dangerous and more difficult to handle.) Under these conditions, high yields of bromohydrins can be obtained. For example,



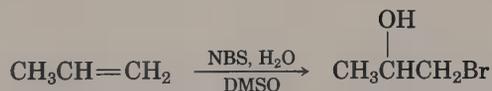
Note that the aromatic ring in the previous example is inert to bromine under the conditions used, even though it contains three carbon-carbon double bonds. Aromatic rings are a good deal more stable than might be expected, a property that will be examined in Chapter 15.

PROBLEM .....

- 7.5 What product(s) would you expect from the reaction of cyclopentene with NBS and water? Show the stereochemistry.

PROBLEM .....

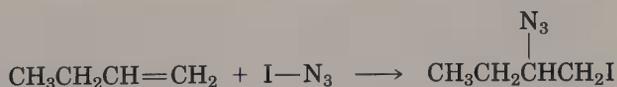
- 7.6 When an unsymmetrical alkene such as propene is treated with *N*-bromosuccinimide in aqueous dimethyl sulfoxide, the major product has the bromine atom bonded to the less substituted carbon atom:



How can you account for this result? Is this Markovnikov or non-Markovnikov orientation?

PROBLEM .....

- 7.7 Iodine azide,  $\text{I}-\text{N}_3$ , adds to alkenes by an electrophilic mechanism similar to that of bromine. If a monosubstituted alkene is used, only one product results:

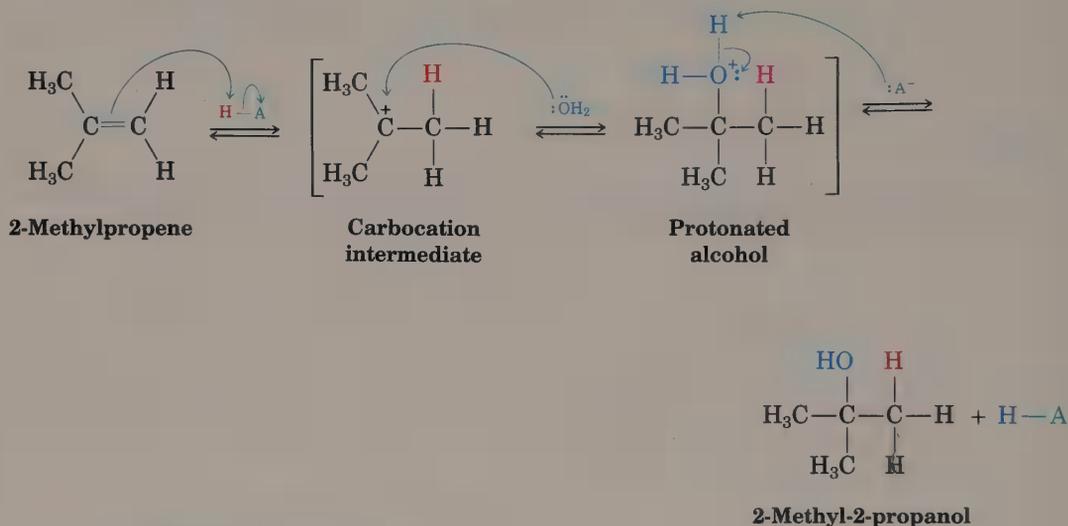


In light of this result, what is the polarity of the  $\text{I}-\text{N}_3$  bond? Propose a mechanism for the reaction.

.....

## 7.4 Hydration of Alkenes: Oxymercuration

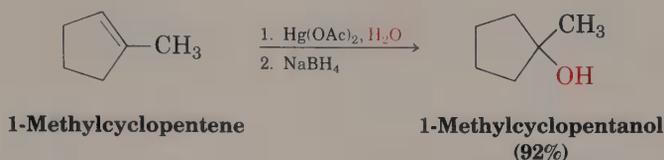
Water can be added to simple alkenes such as ethylene and 2-methylpropene to yield alcohols, ROH, a process called **hydration**. The reaction takes place on treatment of the alkene with water and a strong acid catalyst (HA) by a mechanism similar to that of HX addition. Thus, protonation of an alkene double bond yields a carbocation intermediate that reacts with water as nucleophile to yield a protonated alcohol product (ROH<sub>2</sub><sup>+</sup>). Loss of H<sup>+</sup> from this protonated alcohol gives the neutral alcohol and regenerates the acid catalyst (Figure 7.5).



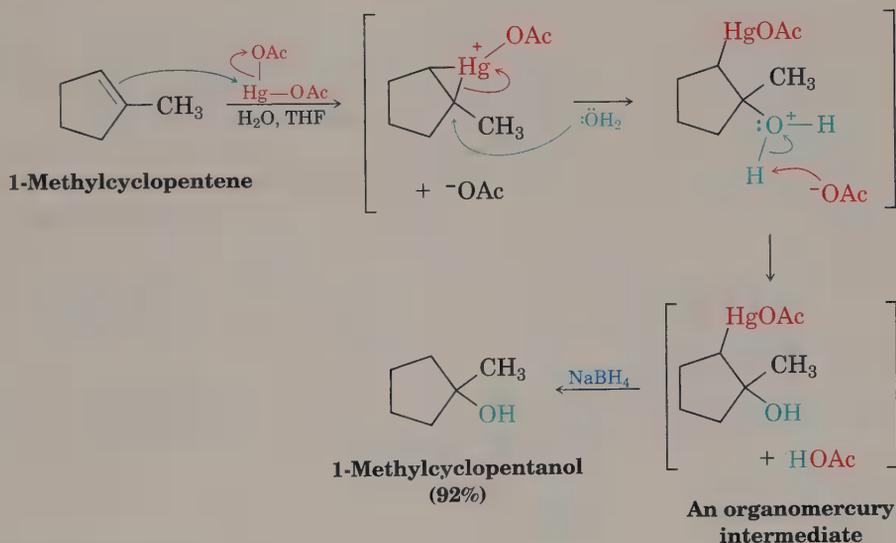
**Figure 7.5** Mechanism of the acid-catalyzed hydration of an alkene.

Acid-catalyzed alkene hydration is suitable for large-scale industrial procedures, and approximately 300,000 tons per year of ethanol are manufactured in the United States by hydration of ethylene at 300°C. The reaction is of little value for most laboratory applications, though, because the high temperatures and strongly acidic conditions required are too vigorous for many organic molecules to survive.

In practice, most alkenes are best hydrated by the **oxymercuration** procedure. When an alkene is treated with mercuric acetate [ $\text{Hg}(\text{O}_2\text{CCH}_3)_2$ , usually abbreviated  $\text{Hg}(\text{OAc})_2$ ] in aqueous tetrahydrofuran (THF) solvent, electrophilic addition to the double bond rapidly occurs. The intermediate *organomercury* compound is then treated with sodium borohydride,  $\text{NaBH}_4$ , and an alcohol is produced. For example:



Alkene oxymercuration is closely analogous to halohydrin formation and is initiated by electrophilic addition of mercuric ion to the alkene to give an intermediate *mercurinium ion*, whose structure resembles that of a bromonium ion (Figure 7.6). Nucleophilic attack of water followed by loss of a proton then yields a stable organomercury addition product. The final step, reaction of the organomercury compound with sodium borohydride, is not well understood but appears to involve radicals as intermediates. Note that the regiochemistry of the reaction corresponds to Markovnikov addition of water; that is, the hydroxyl group becomes attached to the more highly substituted carbon atom, and the hydrogen becomes attached to the less highly substituted carbon.



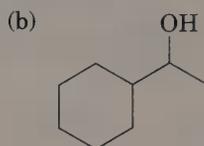
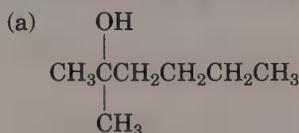
**Figure 7.6** Mechanism of the oxymercuration of alkenes to yield alcohols. This electrophilic addition reaction involves a mercurinium ion intermediate, and its mechanism is similar to that of halohydrin formation.

**PROBLEM** .....

- 7.8 What products would you expect from oxymercuration of these alkenes?  
 (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}=\text{CH}_2$   (b) 2-Methyl-2-pentene

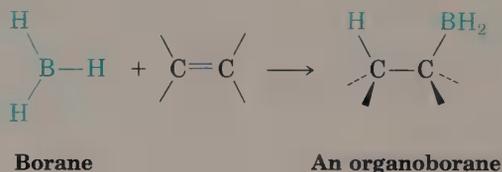
## PROBLEM .....

7.9 What alkenes might these alcohols have been prepared from?

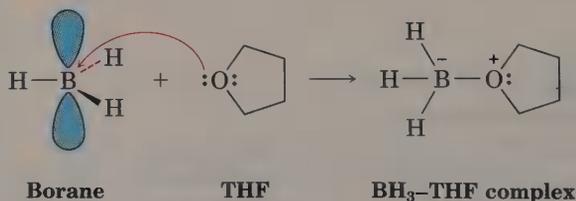


## 7.5 Hydration of Alkenes: Hydroboration

One of the most useful of all alkene additions is the *hydroboration* reaction reported by H. C. Brown<sup>1</sup> in 1959. **Hydroboration** involves addition of a B–H bond of borane,  $\text{BH}_3$ , to an alkene to yield an **organoborane** intermediate,  $\text{RBH}_2$ :



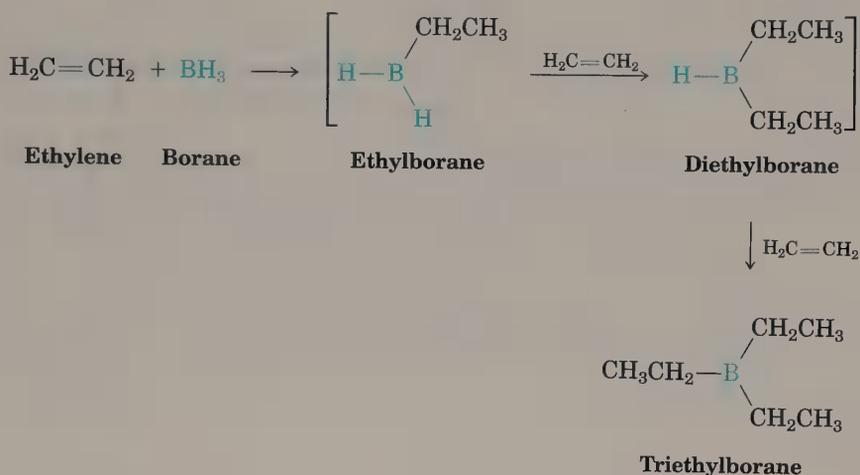
Borane is highly reactive since the boron atom has only six electrons in its valence shell. In tetrahydrofuran (THF) solution, however,  $\text{BH}_3$  accepts an electron pair from a solvent molecule in a Lewis acid–base reaction to complete its octet and form a stable  $\text{BH}_3\text{-THF}$  complex. Although formal-charge calculations for this complex show that there must be a negative charge on boron and a positive charge on oxygen, for all practical purposes the  $\text{BH}_3\text{-THF}$  complex behaves chemically as if it were  $\text{BH}_3$ .



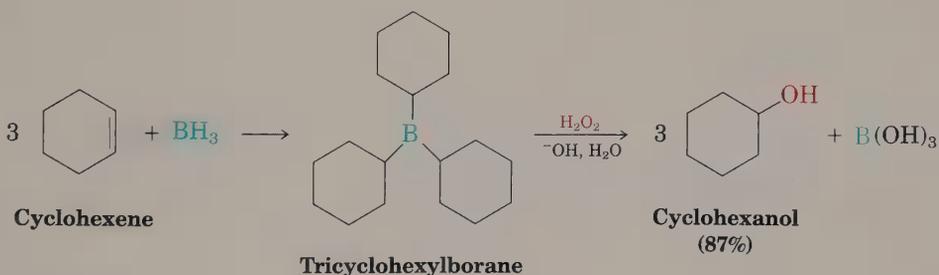
When an alkene reacts with  $\text{BH}_3$  in THF solution, rapid addition to the double bond occurs. Since  $\text{BH}_3$  has three hydrogens, addition occurs three times and a *trialkylborane* product,  $\text{R}_3\text{B}$ , is formed. For example, one molar

<sup>1</sup>Herbert Charles Brown (1912– ); b. London; Ph.D. (1938) University of Chicago (Schlessinger); professor, Purdue University (1947– ); Nobel Prize (1979).

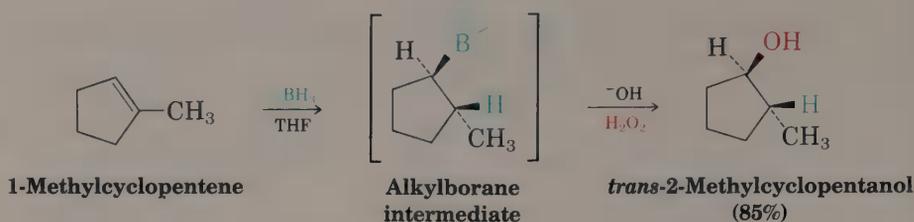
equivalent of  $\text{BH}_3$  adds to three molar equivalents of ethylene to yield triethylborane:



Alkylboranes are extremely useful in synthesis because of the further reactions they undergo. For example, when tricyclohexylborane is treated with aqueous hydrogen peroxide ( $\text{H}_2\text{O}_2$ ) in basic solution, an oxidation takes place. The carbon–boron bond is broken, a hydroxyl group bonds to carbon, and three equivalents of cyclohexanol are produced. The net effect of the two-step hydroboration-plus-oxidation sequence is hydration of the alkene double bond.

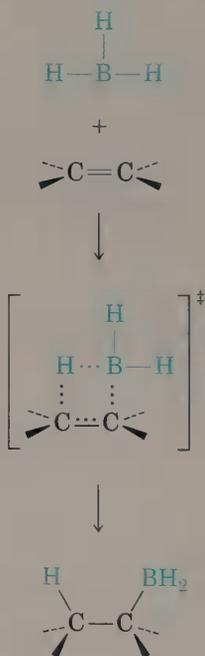


One of the features that makes the hydroboration reaction so useful is the regiochemistry that results when an unsymmetrical alkene is hydroborated. For example, hydroboration–oxidation of 1-methylcyclopentene yields *trans*-2-methylcyclopentanol. Boron and hydrogen add to the alkene with *syn* stereochemistry, with boron attaching to the less highly substituted carbon. During the oxidation step, the boron is replaced by a hydroxyl with the same stereochemistry, resulting in an overall *syn* non-Markovnikov addition of water. This stereochemical result is particularly useful since it is *complementary* to the Markovnikov regiochemistry observed for oxymercuration.



Why does alkene hydroboration take place with “non-Markovnikov” regiochemistry? Hydroboration differs from many other alkene addition reactions in that it occurs in a single step; no carbocation intermediate is involved. We can view the reaction as taking place through a four-center cyclic transition state in which both carbon–hydrogen and carbon–boron bonds form at approximately the same time (Figure 7.7). Since both carbon–hydrogen and carbon–boron bonds form simultaneously from the same face of the alkene, this mechanism explains why *syn* stereochemistry is observed.

Addition of borane to the alkene pi bond occurs in a single step through a cyclic four-membered-ring transition state. The dotted lines indicate partial bonds that are breaking or forming.



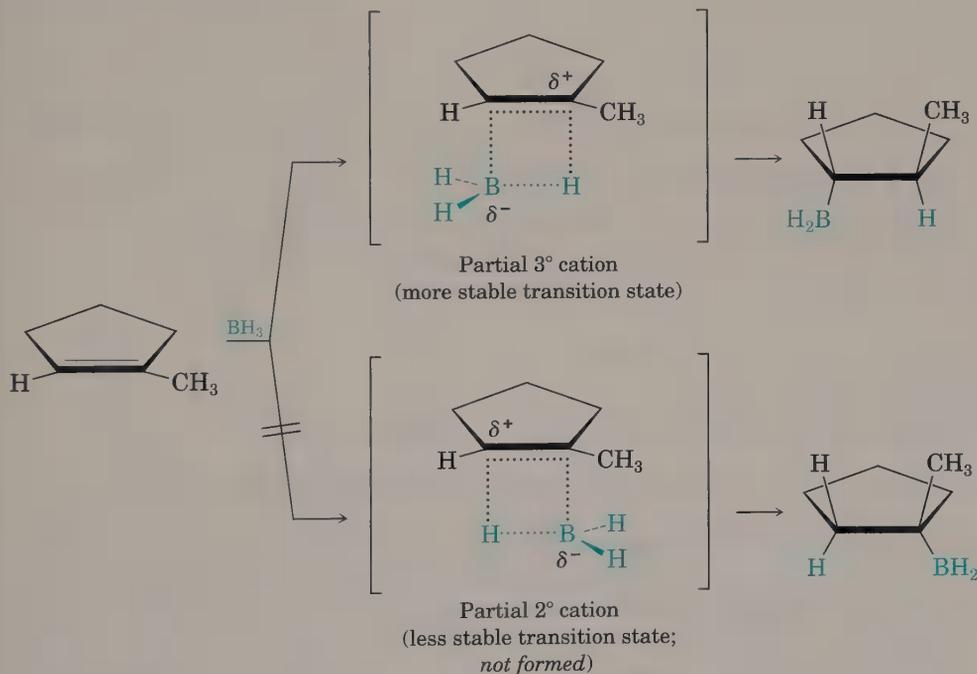
A neutral alkylborane addition product is then formed when reaction is complete.

**Figure 7.7** Mechanism of alkene hydroboration. The reaction occurs in a single step in which both C–H and C–B bonds form at the same time.

The mechanism shown in Figure 7.7 accounts for not only the reaction’s stereochemistry but also its regiochemistry. Although hydroboration does not involve charged intermediates as other alkene addition reactions do, the interaction of borane with an alkene nevertheless has a large amount

of polar character to it. Borane, with only six valence electrons on boron, is a powerful electrophile because of its vacant  $p$  orbital. Thus, the interaction of  $\text{BH}_3$  with a nucleophilic alkene involves a partial transfer of electrons from the alkene to boron, with consequent buildup of polar character in the four-membered-ring transition state. Boron thus carries a partial negative charge ( $\delta^-$ ) since it has gained electrons, and the alkene carbons carry a partial positive charge ( $\delta^+$ ) since they have lost electrons.

In the addition of  $\text{BH}_3$  to an unsymmetrically substituted alkene such as 1-methylcyclopentene, there are two possible four-center transition states. From what we know about the nature of electrophilic additions, we would expect the transition state that places a partial positive charge on the more highly substituted carbon to be favored over the alternative that places the charge on the less highly substituted carbon. Thus, boron tends to add to the less highly substituted carbon (Figure 7.8).

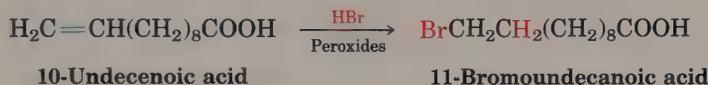


**Figure 7.8** Mechanism of the hydroboration of 1-methylcyclopentene. The favored transition state is the one that places the partial positive charge on the more highly substituted carbon.

In addition to electronic factors, a steric factor is probably also involved in determining the regiochemistry of hydroboration. Attachment of the boron group is favored at the less sterically hindered carbon atom of the alkene, rather than at the more hindered carbon, because there is less steric crowding in the resultant transition state:

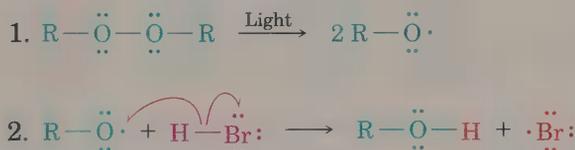


A careful examination of the reaction ultimately resolved the problem when it was found that HBr (but not HCl or HI) can add to alkenes by two entirely different mechanisms. It turns out that 3-bromopropene, the substrate many investigators had been using, is highly air-sensitive and absorbs oxygen to form *peroxides* (R–O–O–R, compounds with oxygen–oxygen bonds). Since the oxygen–oxygen bond is weak and easily broken, peroxides are excellent sources of radicals and can serve to catalyze a *radical addition* of HBr to the alkene, rather than a polar electrophilic addition. In subsequent work, it was found that non-Markovnikov radical addition of HBr occurs with many different alkenes. For example:

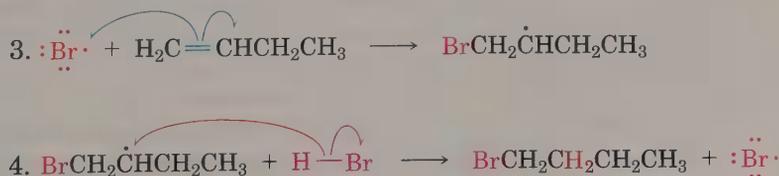


Peroxide-catalyzed addition of HBr to alkenes proceeds by a radical chain mechanism involving addition of bromine radical, Br·, to the alkene. As with the radical chain process we saw earlier for the light-induced chlorination of methane (Section 5.4), both initiation steps and propagation steps are required. In writing the mechanism of this radical reaction, note how a curved half-arrow, or “fishhook”  $\frown$ , is used to show the movement of a single electron, as opposed to the full arrow used to show the movement of an electron pair in a polar reaction.

*Initiation steps* The reaction is initiated in two steps. In the first step, light-induced homolytic cleavage of the weak oxygen–oxygen peroxide bond generates two alkoxy radicals, RO·. An alkoxy radical then abstracts a hydrogen atom from HBr in the second initiation step to give a bromine radical, ·Br.



*Propagation steps* Once a bromine radical has formed in the initiation steps, a chain-reaction cycle of two repeating propagation steps begins. In the first propagation step, bromine radical adds to the alkene double bond, giving an alkyl radical. One electron from Br· and one electron from the pi bond are used to form the new C–Br bond, leaving an unpaired electron on the remaining double-bond carbon. In the second propagation step, this alkyl radical reacts with HBr to yield addition product plus a bromine radical to cycle back into the first propagation step and carry on the chain reaction.





**TABLE 7.1** Bond Dissociation Energies of Different Kinds of C–H Bonds in Alkanes

Bond broken	Radical products	Bond type <sup>a</sup>	Bond dissociation energy, $\Delta H^\circ$ (kcal/mol) (kJ/mol)	
H—CH <sub>3</sub>	H· + CH <sub>3</sub> ·	Methyl	104	435
H—CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	H· + CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> ·	Primary	98	410
$\begin{array}{c} \text{H} \\   \\ \text{CH}_3\text{CHCH}_3 \end{array}$	H· + CH <sub>3</sub> ·CHCH <sub>3</sub>	Secondary	95	397
H—C(CH <sub>3</sub> ) <sub>3</sub>	H· + (CH <sub>3</sub> ) <sub>3</sub> C·	Tertiary	92	385

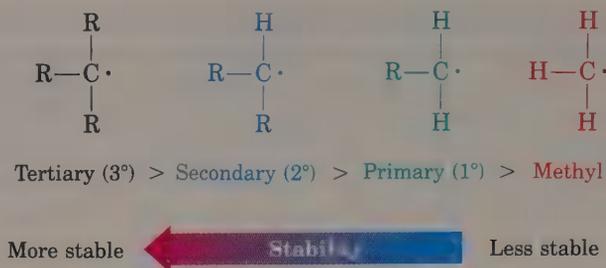
<sup>a</sup>Tertiary C–H bonds are weaker than secondary ones, and secondary bonds are weaker than primary ones.



**Figure 7.10** A plot of C–H bond dissociation energy versus substitution pattern. The order of bond strength is methyl > 1° > 2° > 3°.

both cases (propane), and since one product is the same in both cases (H·), the 3 kcal/mol energy difference is a direct measure of the difference in stability between the primary propyl radical (CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>·) and the secondary propyl radical [(CH<sub>3</sub>)<sub>2</sub>CH·]. In other words, the secondary propyl radical is more stable than the primary propyl radical by 3 kcal/mol.

A similar comparison between a primary C–H bond of 2-methylpropane (98 kcal/mol; 410 kJ/mol) and the tertiary C–H bond in the same molecule (92 kcal/mol; 385 kJ/mol) leads to the conclusion that the tertiary radical is more stable than the primary radical by 6 kcal/mol (25 kJ/mol). We thus find that the stability order of radicals is as follows:



The stability order of radicals is identical to the stability order of carbocations, and for much the same reasons. Carbon radicals, even though they're uncharged, are electron-deficient like carbocations. Thus, the same substituent effects stabilize both.

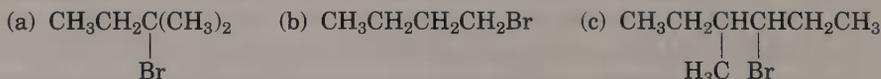
### AN EXPLANATION OF THE REGIOCHEMISTRY OF RADICAL-CATALYZED ADDITION OF HBr TO ALKENES

Now that we know the stability order of radicals, we can complete an explanation for the observed regiochemistry of radical-catalyzed HBr additions to alkenes. Although the addition of bromine radical to an alkene is a slightly exothermic step, there is nevertheless a certain amount of developing radical character in the transition state. Thus, according to the Hammond postulate (Section 6.12), the same factors that make a secondary radical more stable than a primary one also make the transition state leading to the secondary radical more stable than the transition state leading to the primary one. In other words, the more stable radical forms faster than the less stable one.

We conclude that, although peroxide-catalyzed radical addition of HBr to an alkene gives an *apparently* non-Markovnikov product, the reaction in fact proceeds through the more stable intermediate just as a polar reaction does. A reaction energy diagram for the overall process is shown in Figure 7.11.

#### PROBLEM .....

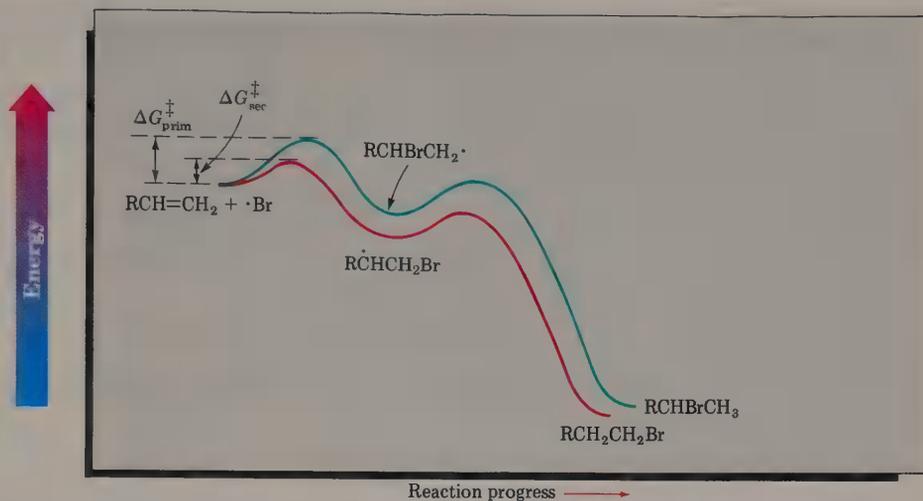
**7.12** Show how you would synthesize the following compounds. Identify the alkene starting material and indicate what reagents you would use in each case.



#### PROBLEM .....

**7.13** Draw a reaction energy diagram for the radical addition of HBr to 2-methyl-2-butene (consider only the propagation steps). Construct your diagram so that reactions leading to the two possible addition products are both shown. Which of the two curves has the lower activation energy for the first step? Which of the two curves has the lower  $\Delta G^\circ$  for the first step?

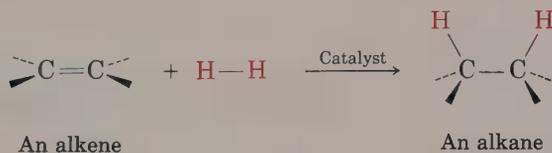
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**Figure 7.11** Reaction energy diagram for the addition of bromine radical to an alkene. The more stable secondary radical forms faster than the less stable primary radical.

## 7.7 Hydrogenation of Alkenes

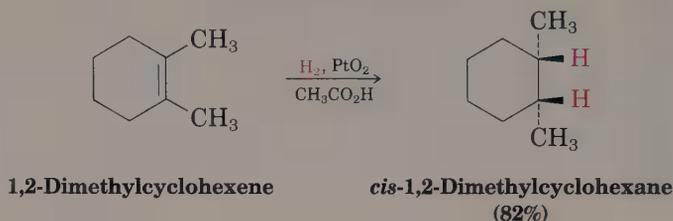
Alkenes react with hydrogen in the presence of a suitable catalyst to yield the corresponding saturated alkane addition products. We describe the result by saying that the double bond has been **hydrogenated**, or *reduced*. Although we looked briefly at catalytic hydrogenation as a method of determining alkene stabilities in Section 6.7, the reaction also has enormous practical value.



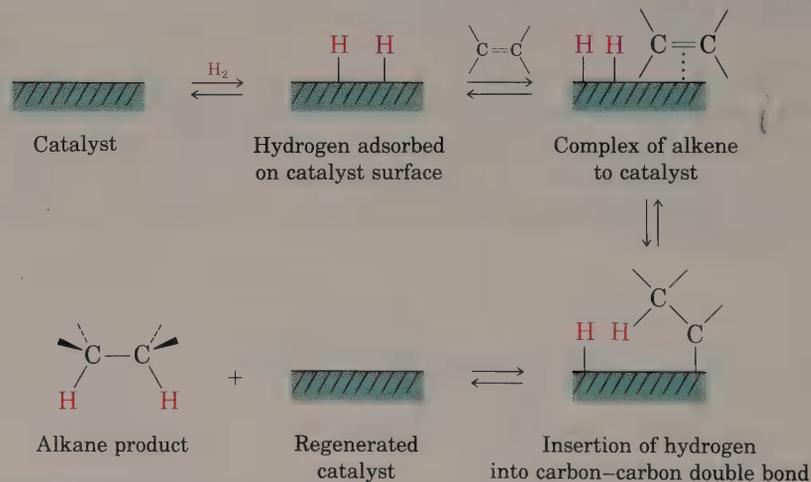
Platinum and palladium are the two catalysts used for most alkene hydrogenations. Palladium is normally employed in a very finely divided state “supported” on an inert material such as charcoal to maximize surface area (Pd/C). Platinum is normally used as PtO<sub>2</sub>, a reagent known as **Adams’ catalyst** after its discoverer, Roger Adams.<sup>2</sup>

<sup>2</sup>Roger Adams (1889–1971); b. Boston; Ph.D. Harvard (Torrey, Richards) (1912); professor, University of Illinois (1916–1971).

Catalytic hydrogenation, unlike most other organic reactions, is a *heterogeneous* process rather than a homogeneous one. That is, the hydrogenation reaction occurs on the surface of solid catalyst particles rather than in solution. For this reason, catalytic hydrogenation has proven difficult to study mechanistically. Observation has shown, however, that hydrogenation usually occurs with *syn* stereochemistry: Both hydrogens add to the double bond from the same face.



The first step in the reaction is adsorption of hydrogen onto the catalyst surface, followed by complexation between catalyst and alkene as vacant metal orbitals overlap with the filled alkene pi orbital. In the final steps, hydrogen is inserted into the double bond, and the saturated product diffuses away from the catalyst (Figure 7.12). It's clear from this picture why the stereochemistry of hydrogenation is *syn*: It *must* be *syn* because both hydrogens add to the double bond from the same catalyst surface.

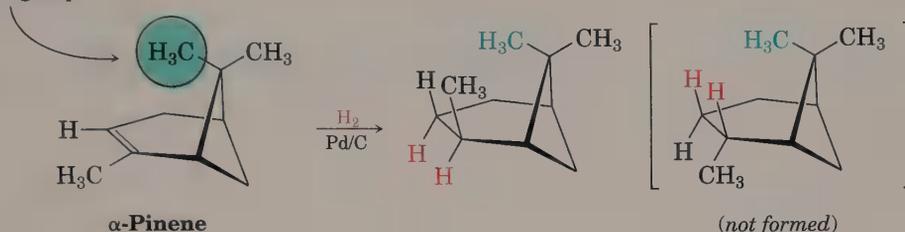


**Figure 7.12** Mechanism of alkene hydrogenation. The reaction takes place with *syn* stereochemistry on the surface of catalyst particles.

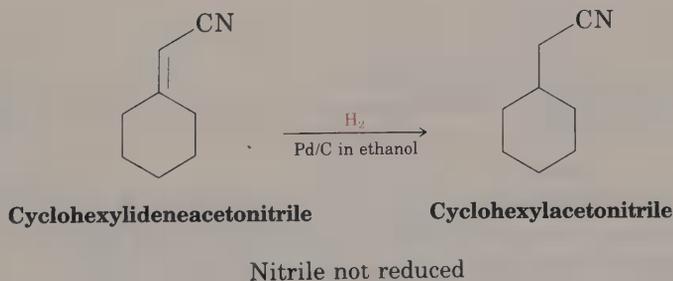
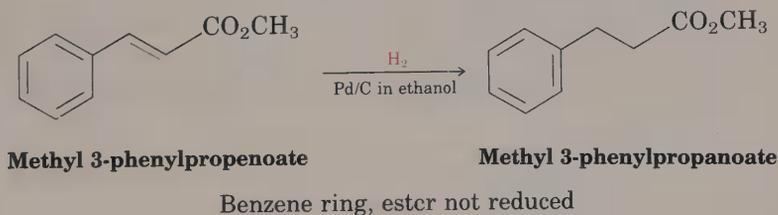
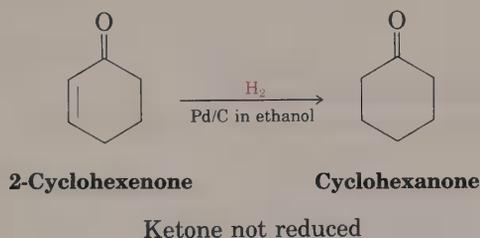
Another stereochemical feature of catalytic hydrogenation is that the reaction is sensitive to the steric environment around the double bond. As a result, the catalyst often approaches only one face, giving rise to a single product. In  $\alpha$ -pinene, for example, one of the methyl groups attached to the

four-membered ring hangs over the top face of the double bond and blocks attack of the hydrogenation catalyst. Reduction therefore occurs exclusively from the bottom face to yield the product shown.

Top side of double bond blocked by methyl group

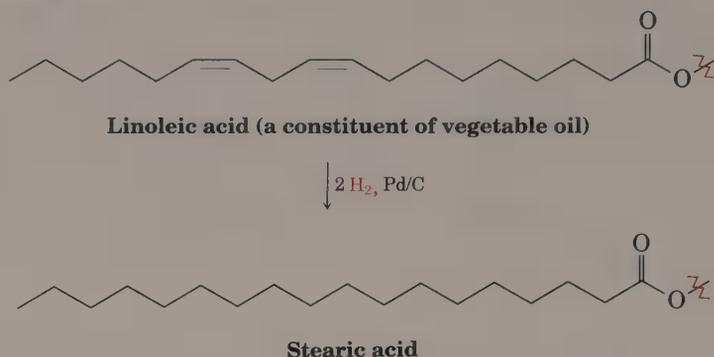


Alkenes are much more reactive than most other functional groups toward catalytic hydrogenation, and the reaction is therefore quite selective. Other functional groups such as ketones, esters, and nitriles survive normal alkene hydrogenation conditions unchanged. Reaction with these groups does occur under more vigorous conditions, however.



Note that, in the hydrogenation of methyl 3-phenylpropenoate, the benzene ring functional group is not affected by hydrogen on palladium even though it contains three double bonds. This is yet another example of the remarkable lack of reactivity of aromatic rings.

In addition to its usefulness in the laboratory, catalytic hydrogenation is of great commercial value in the food industry. Unsaturated vegetable oils, which usually contain numerous double bonds, are catalytically hydrogenated on a vast scale to produce the saturated fats used in margarine and solid cooking fats.

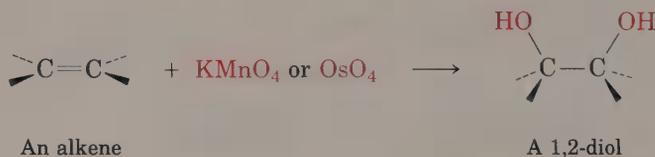


PROBLEM .....

- 7.14 What product would you obtain from catalytic hydrogenation of these alkenes?  
 (a)  $(\text{CH}_3)_2\text{C}=\text{CHCH}_2\text{CH}_3$                       (b) 3,3-Dimethylcyclopentene
- .....

## 7.8 Hydroxylation of Alkenes

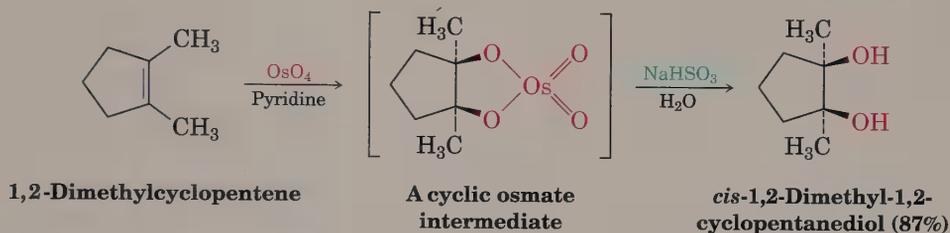
**Hydroxylation** of an alkene—the addition of an  $-\text{OH}$  group to each of the two alkene carbons—can be carried out with reagents such as potassium permanganate ( $\text{KMnO}_4$ ) and osmium tetroxide ( $\text{OsO}_4$ ). Since oxygen is added to the alkene during the reaction, we call this an *oxidation*. Both of these hydroxylation reactions occur with *syn*, rather than *anti*, stereochemistry and yield 1,2-dialcohols, or **diols** (also called **glycols**).



Potassium permanganate is a common and inexpensive oxidant that reacts with alkenes under carefully controlled conditions in alkaline medium to yield 1,2-diols. Unlike many other addition reactions we've studied, hydroxylation does not involve a carbocation intermediate. Instead, the reaction occurs through an intermediate cyclic *manganate* species, formed in a single step by addition of permanganate ion to the alkene. This cyclic manganate is then hydrolyzed to give the diol product. For example, oxidation of cyclohexene, which was first carried out in 1879 by Vladimir Markovnikov, gives *cis*-1,2-cyclohexanediol in rather low yield. This result is typical: *Syn* hydroxylation is always observed, but yields are sometimes poor because of side reactions.

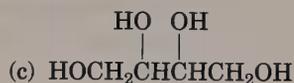
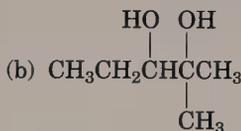
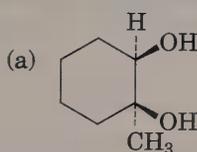


For small-scale laboratory preparations, osmium tetroxide in place of potassium permanganate is much preferred. Although both toxic and expensive,  $\text{OsO}_4$  reacts with alkenes to give high yields of *cis* 1,2-diols. As with permanganate hydroxylation, reaction with  $\text{OsO}_4$  occurs through a cyclic intermediate, which is then cleaved to yield a *cis* diol. A cyclic *osmate* is less reactive than a cyclic manganate, however, and must be cleaved in a separate step. Aqueous sodium bisulfite,  $\text{NaHSO}_3$ , is often used to accomplish this cleavage:



PROBLEM .....

- 7.15 How would you prepare the following compounds? Show the starting alkene and the reagents you would use.



PROBLEM .....

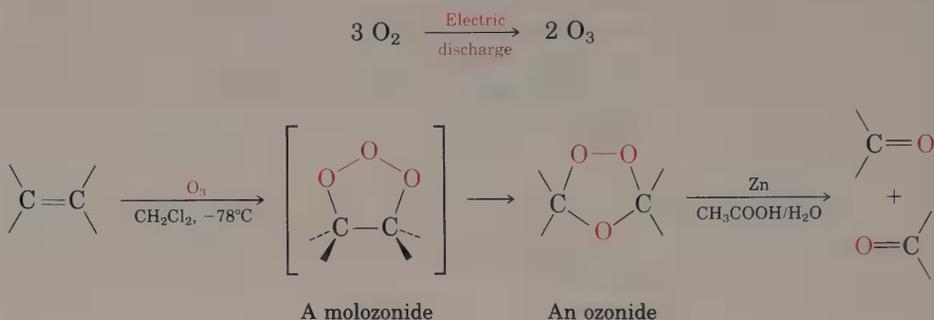
- 7.16 Explain the observation that hydroxylation of *cis*-2-butene with  $\text{OsO}_4$  yields a different product than hydroxylation of *trans*-2-butene. First draw the structure and show the stereochemistry of each product, and then make molecular models. (We'll explore the stereochemistry of the products in more detail in Chapter 9.)
- .....

## 7.9 Oxidative Cleavage of Alkenes

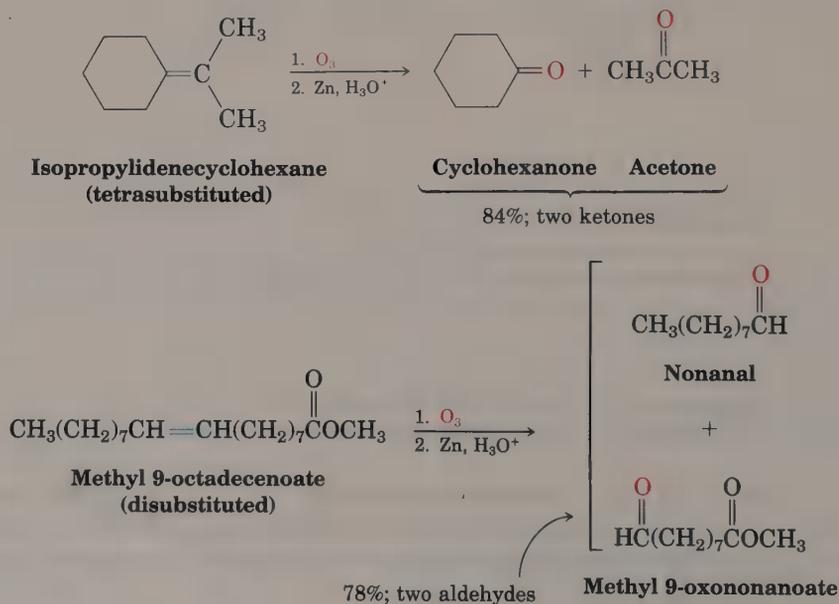
In all of the alkene addition reactions we've seen thus far, the carbon skeleton of the starting material has been left intact. The carbon-carbon double bond has been converted into new functional groups (halide, alcohol, 1,2-diol) by adding different reagents, but no carbon bond framework has been

broken or rearranged. There are, however, powerful oxidizing reagents that *cleave* carbon–carbon double bonds to produce two fragments.

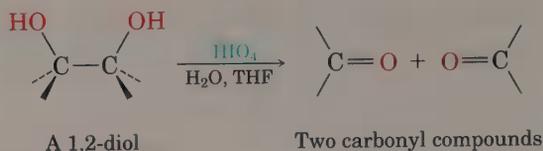
Ozone ( $O_3$ ) is the most useful double-bond cleavage reagent. Prepared conveniently in the laboratory by passing a stream of oxygen through a high-voltage electrical discharge, ozone adds rapidly to alkenes at low temperature to give cyclic intermediates called **molozonides**. Once formed, molozonides then rapidly rearrange to form **ozonides**. (We won't study the mechanism of this unusual rearrangement in detail. It involves the molozonide coming apart into two fragments, which then recombine in a new way.)



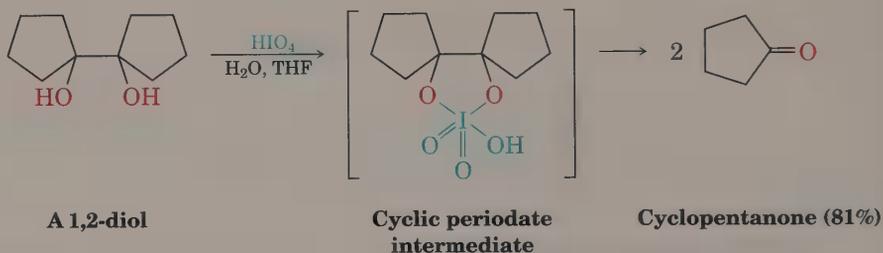
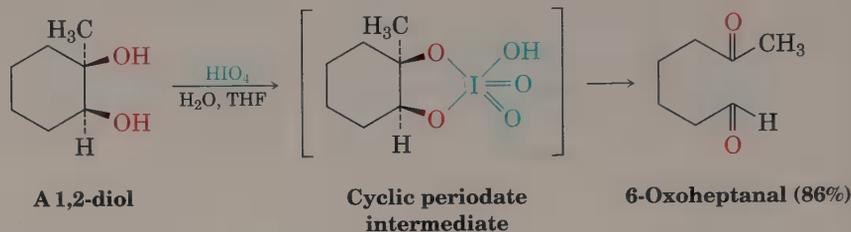
Low-molecular-weight ozonides are explosive and are therefore never isolated. Instead, ozonides are further treated with a reducing agent such as zinc metal in acetic acid to convert them to carbonyl compounds. The net result of the **ozonolysis**–zinc reduction sequence is that the carbon–carbon double bond is cleaved, and oxygen becomes doubly bonded to each of the original alkene carbons. If an alkene with a tetrasubstituted double bond is ozonized, two ketone fragments result; if an alkene with a trisubstituted double bond is ozonized, one ketone and one aldehyde result, and so on.





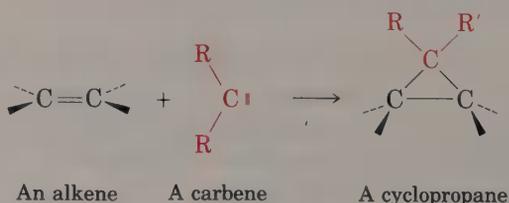


If the two hydroxyls are on an open chain, two carbonyl compounds result. If the two hydroxyls are on a ring, a single dicarbonyl compound is formed. As indicated in the following examples, the cleavage reaction is believed to take place through a cyclic periodate intermediate.



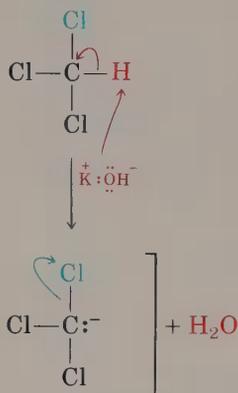
## 7.11 Addition of Carbenes to Alkenes: Cyclopropane Synthesis

The last alkene addition reaction we'll consider is the addition of a carbene to an alkene to yield a cyclopropane. A **carbene**,  $\text{R}_2\text{C}:$ , is a neutral molecule containing a divalent carbon that has only six electrons in its valence shell. It is therefore highly reactive and can be generated only as a reaction intermediate, rather than as an isolable molecule. Because they have only six valence electrons, carbenes are electron-deficient and behave as electrophiles. Thus, they react with nucleophilic carbon-carbon double bonds in a single step without intermediates.



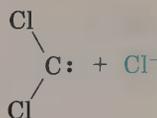
One of the best methods for generating a substituted carbene is by treatment of chloroform,  $\text{CHCl}_3$ , with a strong base such as potassium hydroxide. Loss of a proton from  $\text{CHCl}_3$  generates the trichloromethanide anion,  $^-\text{CCl}_3$ , which expels a chloride ion to give dichlorocarbene,  $:\text{CCl}_2$  (Figure 7.13).

Strong base abstracts the chloroform proton, leaving behind the electron pair from the C–H bond and forming the trichloromethanide anion.



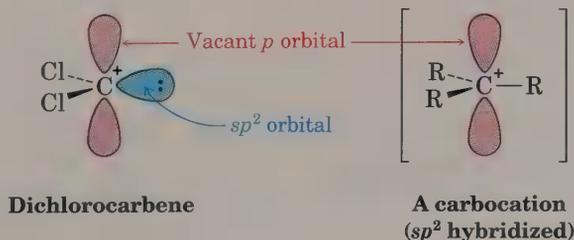
Trichloromethanide anion

Loss of a chloride ion and associated electrons from the C–Cl bond yields the neutral dichlorocarbene.



**Figure 7.13** Mechanism of the formation of dichlorocarbene by reaction of chloroform with strong base.

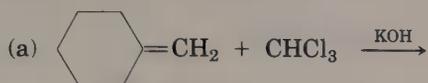
The dichlorocarbene carbon atom is  $sp^2$  hybridized, with a vacant  $p$  orbital extending above and below the plane of the three atoms and with an unshared pair of electrons occupying the third  $sp^2$  lobe. Note that this electronic description of dichlorocarbene is similar to that for carbocations (Section 6.11), with respect both to the  $sp^2$  hybridization of carbon and to the presence of a vacant  $p$  orbital.





## PROBLEM .....

7.19 What products would you expect from these reactions?



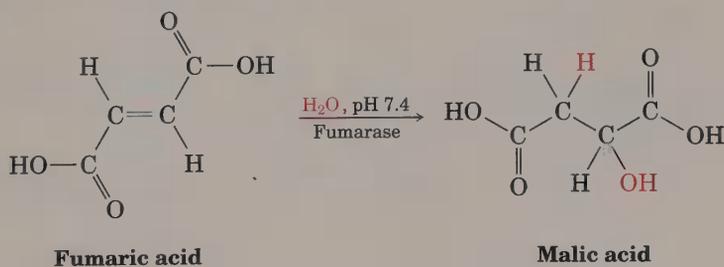
## PROBLEM .....

7.20 Simmons–Smith reaction of cyclohexene with diiodomethane gives a single cyclopropane product, but reaction with 1,1-diiodoethane gives (in low yield) a mixture of two isomeric methylcyclopropane products. Formulate the reactions, and account for the formation of the product mixture.

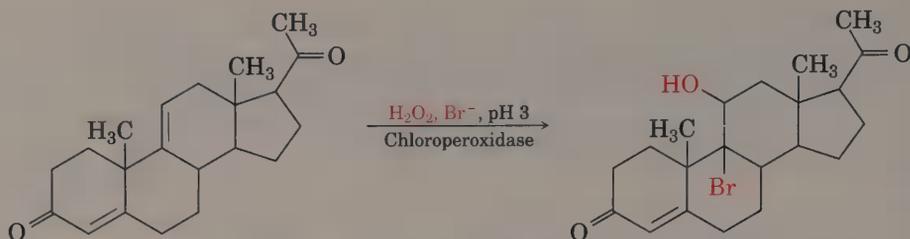
## 7.12 Some Biological Alkene Addition Reactions

The chemistry of living organisms is a fascinating field of study; the simplest one-celled organism is capable of more complex organic synthesis than any human. Yet as we learn more, it becomes clear that the same principles which apply to laboratory chemistry also apply to biochemistry.

Biological organic chemistry takes place in the aqueous medium inside cells rather than in organic solvents and involves complex catalysts called *enzymes*. Nevertheless, the *kinds* of reactions are remarkably similar. Thus, there are many cases of biological addition reactions to alkenes. For example, the enzyme fumarase catalyzes the addition of water to fumaric acid much as sulfuric acid might catalyze the addition of water to ethylene on an industrial scale:



Enzymes can even catalyze the formation of bromohydrins from alkenes. For example, the enzyme chloroperoxidase can catalyze the addition of HO–Br to double bonds:



Enzyme-catalyzed reactions are usually much more chemically selective than their laboratory counterparts. Fumarase, for example, is completely inert toward maleic acid, the *cis* isomer of fumaric acid. Nevertheless, the fundamental processes of organic chemistry are the same in the living cell and in the laboratory.

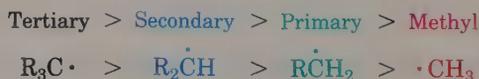
### 7.13 Summary and Key Words

Methods for the preparation of alkenes generally involve **elimination reactions** such as the treatment of an alkyl halide with a strong base to effect elimination of HX (**dehydrohalogenation**). Alcohols undergo a similar elimination of water (**dehydration**) by treatment with strong acid.

**Electrophilic addition** is the most important reaction of alkenes. HCl, HBr, and HI add to carbon–carbon double bonds by a two-step mechanism involving initial reaction of the nucleophilic double bond with  $H^+$  to form a carbocation intermediate, followed by attack of halide ion nucleophile on the cation intermediate. Bromine and chlorine add to alkenes via three-membered-ring **halonium ion** intermediates to give addition products having **anti stereochemistry**. If water is present during halogen addition reactions, a **halohydrin** is formed.

**Hydration** of alkenes (addition of water) is carried out by either of two procedures, depending on the product desired. **Oxymercuration** involves electrophilic addition of mercuric ion to an alkene, followed by trapping of the cation intermediate with water and treatment with  $NaBH_4$ . **Hydroboration** of alkenes involves addition of borane ( $BH_3$ ) followed by oxidation of the intermediate organoborane with alkaline hydrogen peroxide. The two hydration methods are complementary: Oxymercuration gives the product of Markovnikov addition, whereas hydroboration–oxidation gives the product of non-Markovnikov *syn* addition.

HBr (but not HCl or HI) can also add to alkenes by a radical chain pathway to give the non-Markovnikov product. Radicals have the stability order



Alkenes are **reduced** by addition of hydrogen in the presence of a catalyst such as platinum or palladium. **Catalytic hydrogenation** is a

heterogeneous process that occurs on the surface of catalyst particles rather than in solution, and occurs with syn stereochemistry. Cis 1,2-diols can be made directly from alkenes by **hydroxylation** with osmium tetroxide,  $\text{OsO}_4$ . Alkenes can also be cleaved to produce carbonyl compounds by reaction with ozone, followed by reduction with zinc metal.

**Carbenes,  $\text{R}_2\text{C}:$** , are neutral molecules containing a divalent carbon with only six valence electrons. Carbenes are highly reactive toward alkenes, adding to give cyclopropanes. Dichlorocarbene, usually prepared from  $\text{CHCl}_3$  by reaction with base, adds to alkenes to give 1,1-dichlorocyclopropanes. Nonhalogenated cyclopropanes are best prepared by treatment of the alkene with  $\text{CH}_2\text{I}_2$  and zinc-copper alloy—the **Simmons-Smith reaction**.

## LEARNING REACTIONS

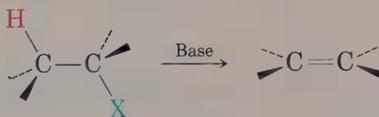
What's seven times nine? Sixty-three, of course. You didn't have to stop and figure it out; you didn't have to count your fingers and toes; you knew the answer immediately because you've memorized the multiplication tables. Learning the reactions of organic chemistry requires the same approach: Reactions have to be memorized for immediate recall if they are to be useful.

Different people take different approaches to learning reactions. Some people make flashcards; others find studying with friends to be helpful. To help guide your study, the chapters in this book end with a summary of the new reactions just presented. In addition, the accompanying *Study Guide and Solutions Manual* has several appendixes that organize organic reactions from other viewpoints. Fundamentally, though, there are no shortcuts. Learning organic chemistry takes effort.

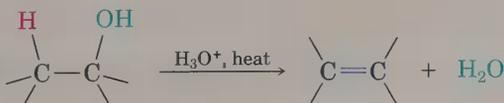
## 7.14 Summary of Reactions

*Note:* No stereochemistry is implied unless specifically indicated with wedged, solid, and dashed lines.

1. Synthesis of alkenes
  - a. Dehydrohalogenation of alkyl halides (Section 7.1)

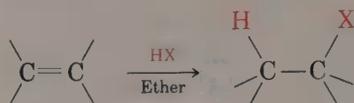


- b. Dehydration of alcohols (Section 7.1)



## 2. Addition reactions of alkenes

- a. Addition of HX, where X = Cl, Br, or I (Sections 6.9 and 6.10)



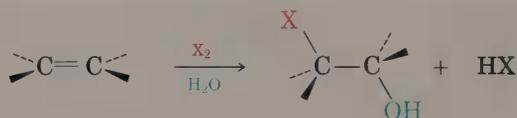
Markovnikov regiochemistry is observed: H adds to the less substituted carbon and X adds to the more substituted carbon.

- b. Addition of halogens, where X
- <sub>2</sub>
- = Cl
- <sub>2</sub>
- or Br
- <sub>2</sub>
- (Section 7.2)



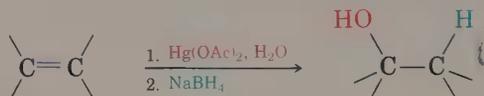
Anti addition is observed.

- c. Halohydrin formation (Section 7.3)



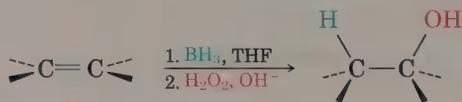
Markovnikov regiochemistry and anti stereochemistry are observed.

- d. Addition of water by oxymercuration (Section 7.4)



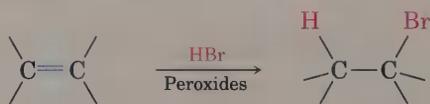
Markovnikov regiochemistry is observed, with the OH attaching to the more substituted carbon.

- e. Addition of water by hydroboration-oxidation (Section 7.5)



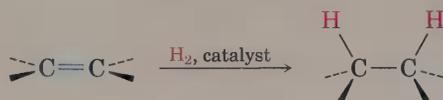
Non-Markovnikov syn addition is observed.

- f. Radical addition of HBr to alkenes (Section 7.6)



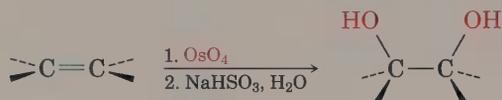
Non-Markovnikov addition is observed.

## g. Hydrogenation of alkenes (Section 7.7)



Syn addition is observed.

## h. Hydroxylation of alkenes (Section 7.8)



Syn addition is observed.

## i. Addition of carbenes to alkenes to yield cyclopropanes (Section 7.11)

## (1) Dichlorocarbene addition

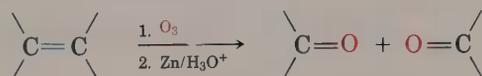
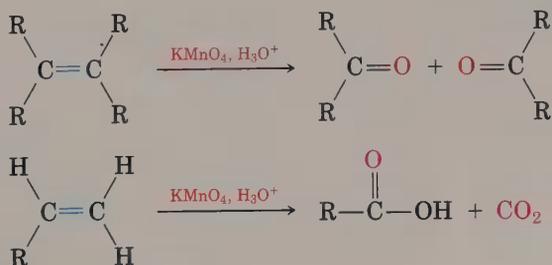


## (2) Simmons–Smith reaction

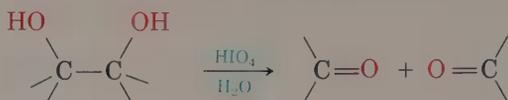


## 3. Oxidative cleavage of alkenes (Section 7.9)

## a. Treatment with ozone, followed by zinc in acetic acid

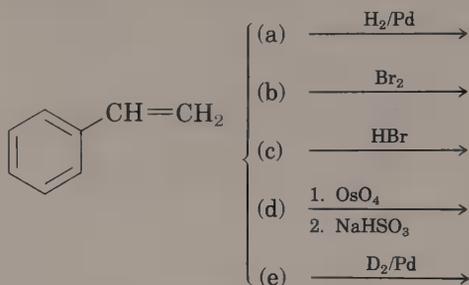
b. Reaction with  $\text{KMnO}_4$  in acidic solution

## 4. Oxidative cleavage of 1,2-diols (Section 7.10)



## ADDITIONAL PROBLEMS .....

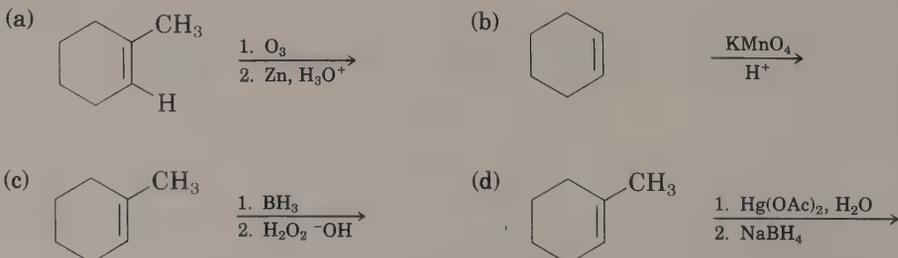
7.21 Predict the products of the following reactions (the benzene ring is unreactive in all cases). Indicate regiochemistry when relevant.



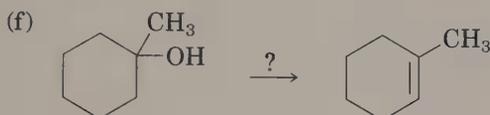
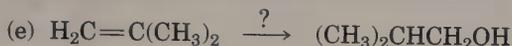
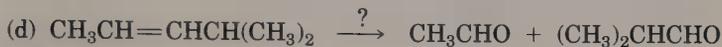
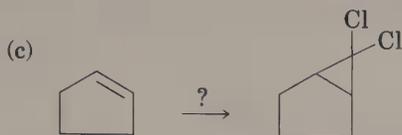
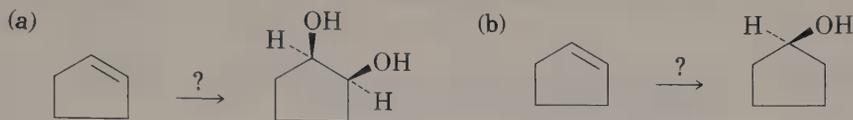
7.22 Suggest structures for alkenes that give the following reaction products. There may be more than one answer for some cases.

- (a) ?  $\xrightarrow{\text{H}_2/\text{Pd}}$  2-Methylhexane
- (b) ?  $\xrightarrow{\text{H}_2/\text{Pd}}$  1,1-Dimethylcyclohexane
- (c) ?  $\xrightarrow{\text{Br}_2/\text{CCl}_4}$  2,3-Dibromo-5-methylhexane
- (d) ?  $\xrightarrow[2. \text{NaBH}_4]{1. \text{Hg}(\text{OAc})_2, \text{H}_2\text{O}}$   $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{OH})\text{CH}_3$
- (e) ?  $\xrightarrow{\text{HBr/peroxides}}$  2-Bromo-3-methylheptane
- (f) ?  $\xrightarrow{\text{HCl, ether}}$  2-Chloro-3-methylheptane

7.23 Predict the products of the following reactions, indicating both regiochemistry and stereochemistry where appropriate.



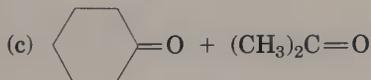
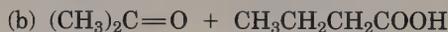
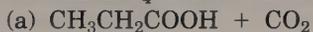
7.24 How would you carry out the following transformations? Indicate the proper reagents.



7.25 Give the structure of an alkene that provides only  $(\text{CH}_3)_2\text{C}=\text{O}$  on ozonolysis followed by treatment with zinc.

7.26 Draw the structure of a hydrocarbon that reacts with 1 mol equiv of hydrogen on catalytic hydrogenation and gives only pentanal,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CHO}$ , on ozonolysis. Write the reactions involved.

7.27 Show the structures of alkenes that give the following products on oxidative cleavage with  $\text{KMnO}_4$  in acidic solution.



7.28 Compound A has the formula  $\text{C}_{10}\text{H}_{16}$ . On catalytic hydrogenation over palladium, it reacts with only 1 equiv of hydrogen. Compound A undergoes reaction with ozone, followed by zinc treatment, to yield a symmetrical diketone, B ( $\text{C}_{10}\text{H}_{16}\text{O}_2$ ).

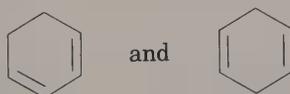
(a) How many rings does A have?

(b) What are the structures of A and B?

(c) Formulate the reactions.

7.29 An unknown hydrocarbon, A, with formula  $\text{C}_6\text{H}_{12}$ , reacts with 1 equiv of hydrogen over a palladium catalyst. Hydrocarbon A also reacts with  $\text{OsO}_4$  to give a diol, B. When oxidized with  $\text{KMnO}_4$  in acidic solution, A gives two fragments. One fragment can be identified as propanoic acid,  $\text{CH}_3\text{CH}_2\text{COOH}$ , and the other fragment can be shown to be a ketone, C. What are the structures of A, B, and C? Write all reactions, and show your reasoning.

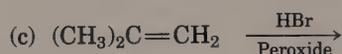
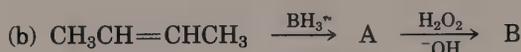
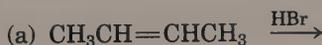
7.30 Using an oxidative cleavage reaction, explain how you would distinguish between the following two isomeric dienes:



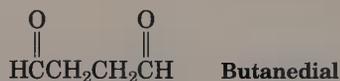
7.31 Compound A,  $C_{10}H_{18}O$ , undergoes reaction with dilute  $H_2SO_4$  at  $250^\circ C$  to yield a mixture of two alkenes,  $C_{10}H_{16}$ . The major alkene product, B, gives only cyclopentanone after ozone treatment followed by reduction with zinc in acetic acid. Formulate the reactions involved, and identify A and B.

7.32 Which reaction would you expect to occur faster, addition of HBr to cyclohexene or to 1-methylcyclohexene? Explain your answer.

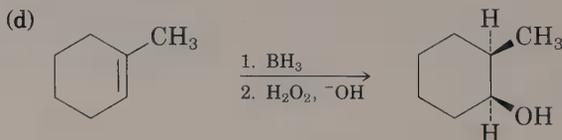
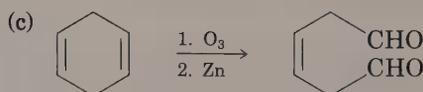
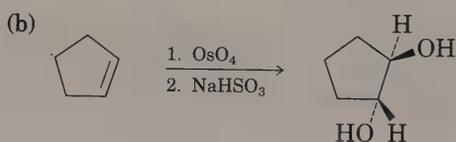
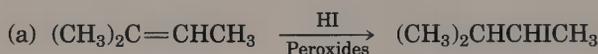
7.33 Predict the products of the following reactions, and indicate regiochemistry if relevant.



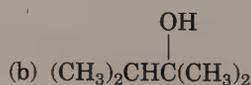
7.34 Draw the structure of a hydrocarbon that absorbs 2 mol equiv of hydrogen on catalytic hydrogenation and gives only butanedial on ozonolysis.

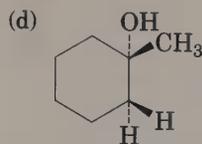
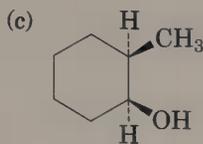


7.35 In planning the synthesis of one compound from another, it's just as important to know what *not* to do as to know what to do. The following proposed reactions all have serious drawbacks to them. Explain the potential problems of each reaction.

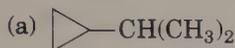


7.36 Which of these alcohols could *not* be made selectively by hydroboration-oxidation of an alkene?

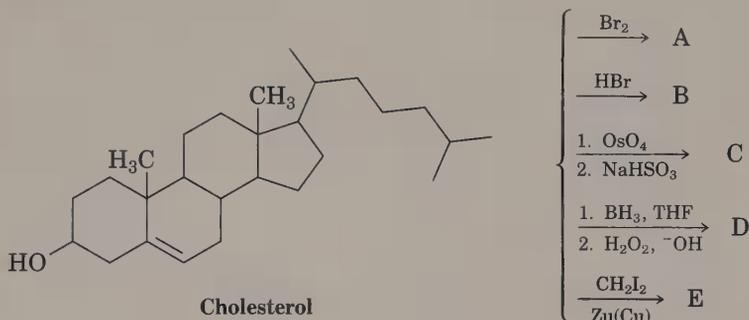




7.37 What alkenes might be used to prepare these cyclopropanes?



7.38 Predict the products of the following reactions. Don't worry about the size of the molecule; concentrate on the functional groups.

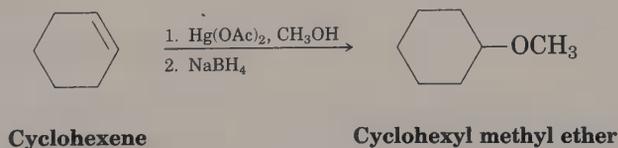


7.39 The sex attractant of the common housefly is a hydrocarbon with the formula  $\text{C}_{23}\text{H}_{46}$ . On treatment with aqueous acidic  $\text{KMnO}_4$ , two products are obtained,  $\text{CH}_3(\text{CH}_2)_{12}\text{COOH}$  and  $\text{CH}_3(\text{CH}_2)_7\text{COOH}$ . Propose a structure.

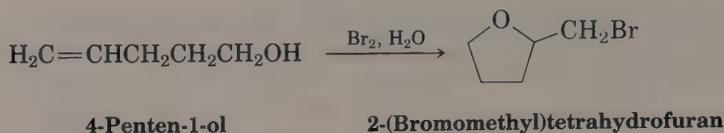
7.40 Compound A has the formula  $\text{C}_8\text{H}_8$ . It reacts rapidly with  $\text{KMnO}_4$  to give  $\text{CO}_2$  and a carboxylic acid, B ( $\text{C}_7\text{H}_6\text{O}_2$ ), but reacts with only 1 equiv of  $\text{H}_2$  on catalytic hydrogenation over a palladium catalyst. On hydrogenation under conditions that reduce aromatic rings, 4 equiv of  $\text{H}_2$  are taken up, and hydrocarbon C ( $\text{C}_8\text{H}_{16}$ ) is produced. What are the structures of A, B, and C? Formulate the reactions.

7.41 Reaction of 2-methylpropene with methyl alcohol in the presence of sulfuric acid catalyst yields methyl *tert*-butyl ether,  $\text{CH}_3\text{OC}(\text{CH}_3)_3$ , by a mechanism analogous to that of acid-catalyzed alkene hydration. Write the mechanism in detail.

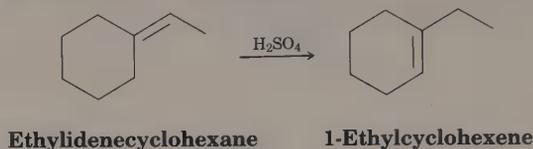
7.42 Reaction of cyclohexene with mercuric acetate in methyl alcohol rather than water, followed by treatment with  $\text{NaBH}_4$ , yields cyclohexyl methyl ether rather than cyclohexanol. Suggest a mechanism to account for this ether synthesis.



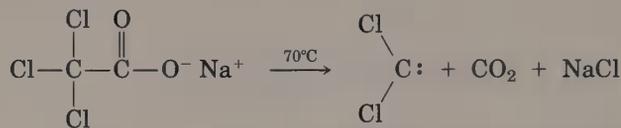
7.43 When 4-penten-1-ol is treated with aqueous bromine, a cyclic bromo ether is formed, rather than the expected bromohydrin. Propose a mechanism for this transformation. [Hint: See Problem 7.42.]



- 7.44 How would you distinguish between the following pairs of compounds using simple chemical tests? Tell what you would do and what you would see.
- (a) Cyclopentene and cyclopentane                      (b) 2-Hexene and benzene
- 7.45 Ethylidenecyclohexane, on treatment with a strong acid, isomerizes to yield 1-ethylcyclohexene. Propose a mechanism for this reaction. Which is the more stable alkene, ethylidenecyclohexane or 1-ethylcyclohexene?



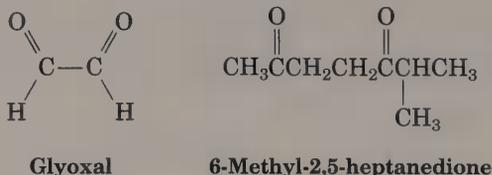
- 7.46 In addition to its preparation from  $\text{CHCl}_3$ , dichlorocarbene can also be generated by heating sodium trichloroacetate:



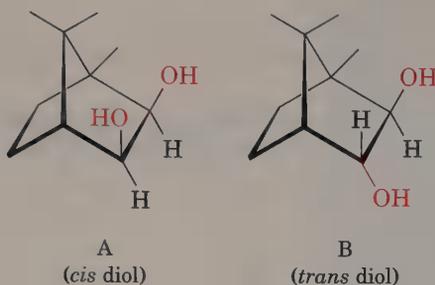
Propose a mechanism to account for this reaction, and use curved arrows to indicate the movement of electrons. What relation does your mechanism bear to the base-induced elimination of HCl from chloroform?

- 7.47  $\alpha$ -Terpinene,  $\text{C}_{10}\text{H}_{16}$ , is a pleasant-smelling hydrocarbon that has been isolated from oil of marjoram. On hydrogenation over a palladium catalyst,  $\alpha$ -terpinene reacts with 2 mol equiv of hydrogen to yield a new hydrocarbon,  $\text{C}_{10}\text{H}_{20}$ . On ozonolysis, followed by reduction with zinc and acetic acid,  $\alpha$ -terpinene yields two products, glyoxal and 6-methyl-2,5-heptanedione.

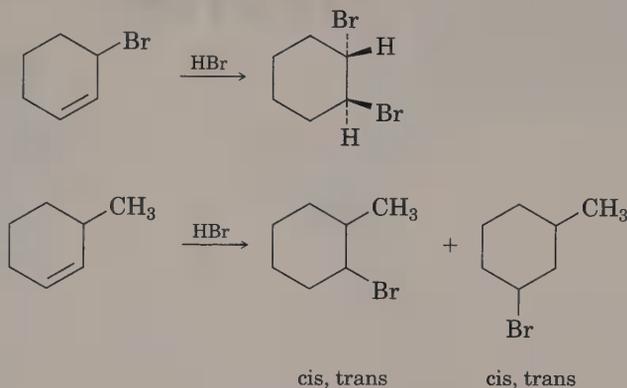
How many degrees of unsaturation does  $\alpha$ -terpinene have? How many double bonds? How many rings? Propose a structure for  $\alpha$ -terpinene that is consistent with the foregoing data.



- 7.48 Evidence that cleavage of 1,2-diols by  $\text{HIO}_4$  occurs through a five-membered cyclic periodate intermediate is based on *kinetic data*: the measurement of reaction rates. When diols A and B were prepared and the rates of their reaction with  $\text{HIO}_4$  were measured, it was found that diol A cleaved approximately 1 million times faster than diol B. Explain these results by making molecular models of A and B and of potential cyclic periodate intermediates.

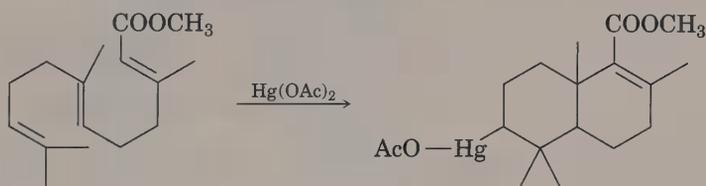


- 7.49 Reaction of HBr with 3-bromocyclohexene yields *trans*-1,2-dibromocyclohexane as the sole product, but reaction of 3-methylcyclohexene with HBr yields a mixture of four products: *cis*- and *trans*-1-bromo-3-methylcyclohexane and *cis*- and *trans*-1-bromo-2-methylcyclohexane.



Draw structures of the possible intermediates, and then explain why only a single product is formed in the reaction of HBr with 3-bromocyclohexene.

- 7.50 The following reaction takes place in high yield:

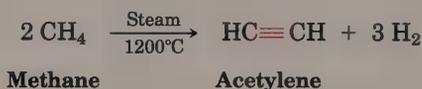


Even though you've never seen the reaction before, use your general knowledge of alkene chemistry to propose a mechanism that accounts for the observed result.

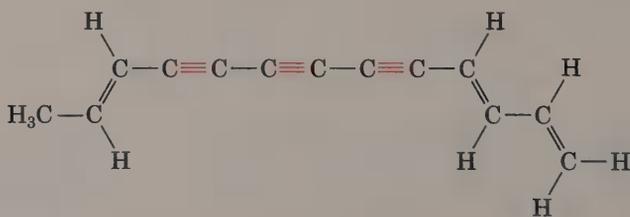
# Chapter 8

## Alkynes

**Alkynes**, also called **acetylenes**, are hydrocarbons that contain a carbon-carbon triple bond. Acetylene itself,  $\text{H}-\text{C}\equiv\text{C}-\text{H}$ , the simplest alkyne, was once widely used in industry as the starting material for the preparation of acetaldehyde, acetic acid, vinyl chloride, and other high-volume chemicals, but more efficient routes using ethylene as starting material are now more common. Acetylene is still used in the preparation of acrylic polymers, however, and is prepared industrially by high-temperature decomposition (*pyrolysis*) of methane. This method is not of general utility in the laboratory.



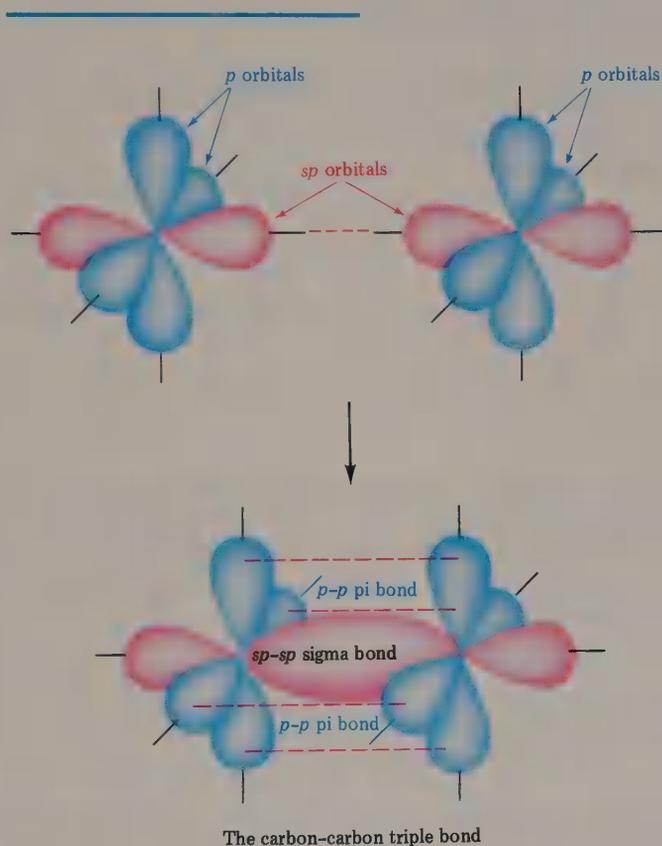
A large number of naturally occurring acetylenic compounds have been isolated from the plant kingdom. For example, the following *triyne* from the safflower, *Carthamus tinctorius* L., evidently forms part of the plant's chemical defenses against nematode infestation:



A triyne from safflower

## 8.1 Electronic Structure of Alkynes

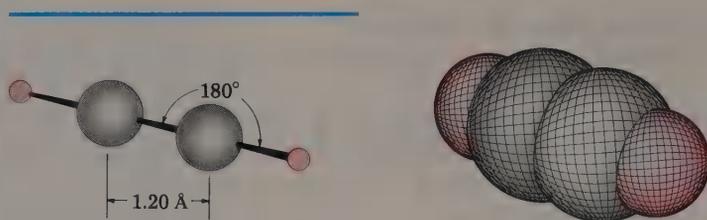
A carbon–carbon triple bond results from the overlap of two  $sp$ -hybridized carbon atoms (Section 1.11). Recall that the two  $sp$ -hybrid orbitals of carbon lie at an angle of  $180^\circ$  to each other along an axis perpendicular to the axes of the two unhybridized  $2p_y$  and  $2p_z$  orbitals. When two  $sp$ -hybridized carbons approach each other for bonding, the geometry is perfect for the formation of one  $sp$ – $sp$  sigma bond and two  $p$ – $p$  pi bonds—a net *triple* bond (Figure 8.1).



**Figure 8.1** Formation of a carbon–carbon triple bond by overlap of two  $sp$ -hybridized carbons.

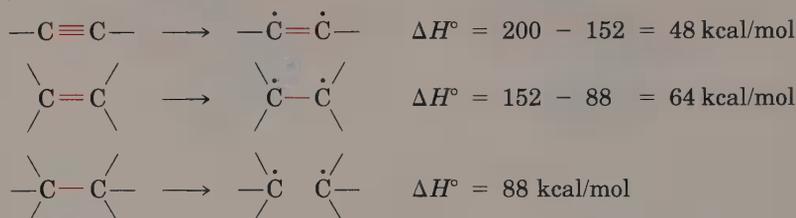
The two remaining  $sp$  orbitals form bonds to other atoms at an angle of  $180^\circ$  from the carbon–carbon bond. Thus, acetylene,  $C_2H_2$ , is a linear molecule with H–C–C bond angles of  $180^\circ$  (Figure 8.2).

The bond length of the carbon–carbon triple bond in acetylene is  $1.20 \text{ \AA}$ , and its strength is approximately  $200 \text{ kcal/mol}$  ( $837 \text{ kJ/mol}$ ), making it the



**Figure 8.2** The structure of acetylene,  $\text{H}-\text{C}\equiv\text{C}-\text{H}$ . The  $\text{H}-\text{C}-\text{C}$  bond angles are  $180^\circ$ .

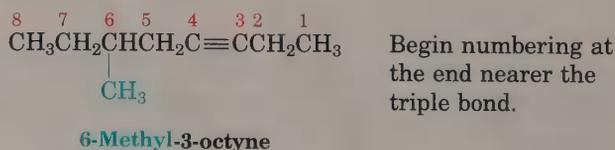
shortest and strongest known carbon-carbon bond. In a purely bookkeeping sense, we can assign bond strengths to each of the three triple-bond “parts.” Since we know that the carbon-carbon *single* bond in ethane has a strength of 88 kcal/mol (368 kJ/mol) and the carbon-carbon *double* bond of ethylene has a strength of 152 kcal/mol (636 kJ/mol), we can “dissect” the overall carbon-carbon triple bond:



Although the two alkyne pi bonds are actually equivalent, this crude calculation shows that approximately 48 kcal/mol (200 kJ/mol) is needed to break the first of them. Since this value is  $64 - 48 = 16$  kcal/mol less than the energy required to break an *alkene* pi bond, we might predict that alkynes should be highly reactive.

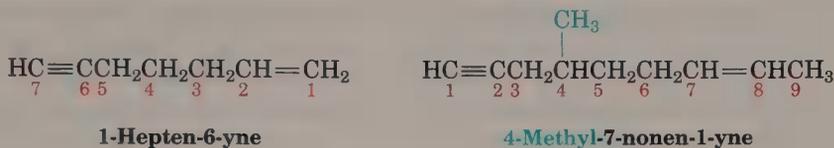
## 8.2 Nomenclature of Alkynes

Alkynes follow the general rules of hydrocarbon nomenclature already discussed. The suffix *-yne* is used in the base hydrocarbon name to denote an alkyne, and the position of the triple bond is indicated by its number in the chain. Numbering always begins at the chain end nearer the triple bond so that the triple bond receives as low a number as possible.

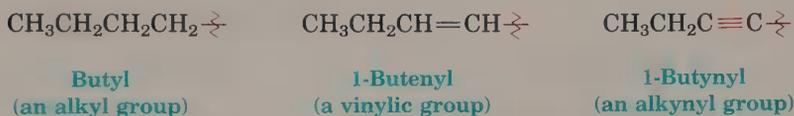


Compounds with more than one triple bond are called *dienes*, *trienes*, and so forth; compounds containing both double and triple bonds are called

*enynes* (not *ynenes*). Numbering of an enyne chain always starts from the end nearer the first multiple bond, whether double or triple. When there is a choice in numbering, though, double bonds receive lower numbers than do triple bonds. For example:

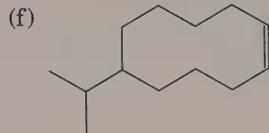
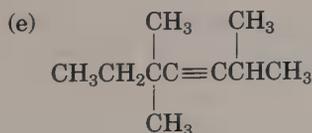
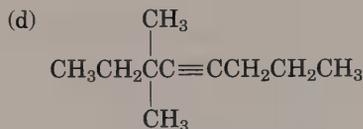
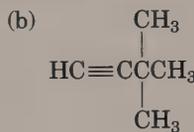
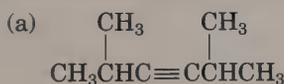


As was the case with hydrocarbon substituents derived from alkanes and alkenes, *alkynyl* groups are also possible:



PROBLEM .....

8.1 Give correct IUPAC names for these compounds:



PROBLEM .....

8.2 There are seven isomeric alkynes with the formula  $\text{C}_6\text{H}_{10}$ . Draw them and name them according to IUPAC rules.

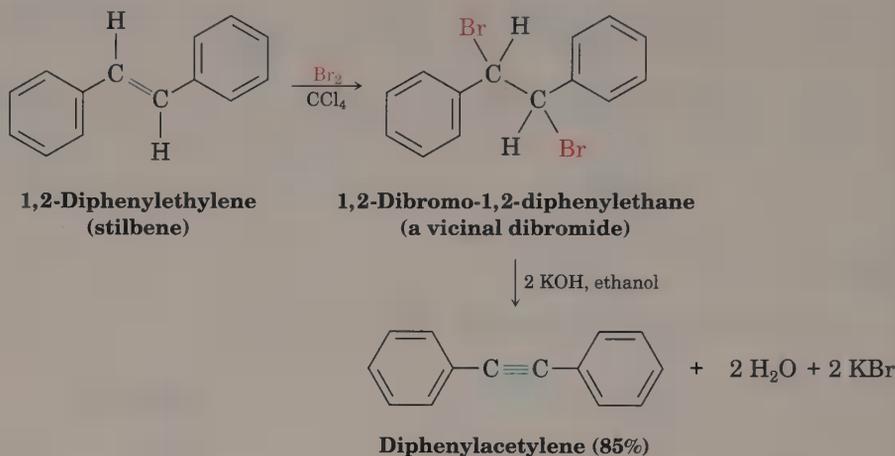
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## 8.3 Preparation of Alkynes: Elimination Reactions of Dihalides

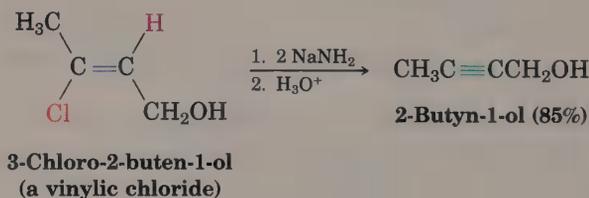
Alkynes can be prepared by elimination of HX from alkyl halides in much the same manner as alkenes can (Section 7.1). Since an alkyne is doubly unsaturated, however, we have to eliminate *two* molecules of HX. Treatment

of a 1,2-dihalide (a **vicinal** dihalide) with excess strong base such as KOH or  $\text{NaNH}_2$  results in a twofold elimination of HX and formation of an alkyne. As with the elimination of HX to form an alkene, we'll defer a discussion of mechanism until Chapter 11.

The necessary vicinal dihalides are themselves readily available by addition of bromine or chlorine to alkenes. Thus, the overall sequence of halogenation–dehydrohalogenation provides an excellent method for going from an alkene to an alkyne. For example, diphenylethylene is converted into diphenylacetylene by reaction with bromine and subsequent base treatment:



Although different bases can be used for dehydrohalogenation, sodium amide ( $\text{NaNH}_2$ ) is normally preferred because it usually gives higher yields. The twofold dehydrohalogenation takes place in discrete steps through a vinylic halide intermediate, which suggests that vinylic halides themselves should give alkynes when treated with strong base. This is indeed the case. For example:

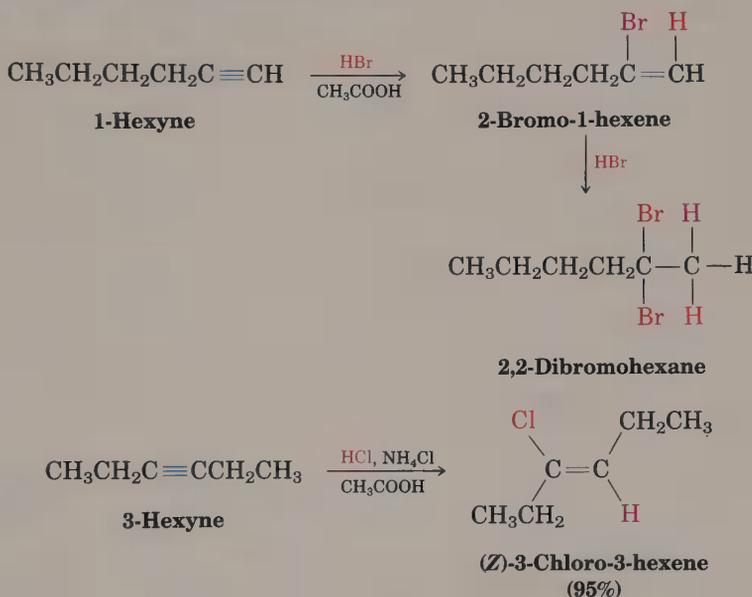


## 8.4 Reactions of Alkynes: Addition of HX and $\text{X}_2$

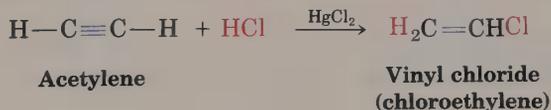
Based on the electronic similarity of alkenes and alkynes, you might expect that the chemical reactivity of the two functional groups should also be similar. The pi part of the triple bond is weak (48 kcal/mol), and the electrons are readily accessible to attacking reagents. Alkynes do indeed exhibit

much chemistry similar to that of alkenes, but there are also significant differences.

As a general rule, electrophilic reagents add to alkynes in the same way that they add to alkenes. With HX, for example, alkynes give the expected addition products. Although the reactions usually can be stopped after addition of one equivalent of HX, an excess of acid leads to a dihalide product. For example, reaction of 1-hexyne with two equivalents of HBr yields 2,2-dibromohexane. As the following examples indicate, the regiochemistry of addition follows Markovnikov's rule. Halogen adds to the more highly substituted side of the alkyne bond, and hydrogen adds to the less highly substituted side. Trans stereochemistry of H and X is normally (though not always) found in the product.



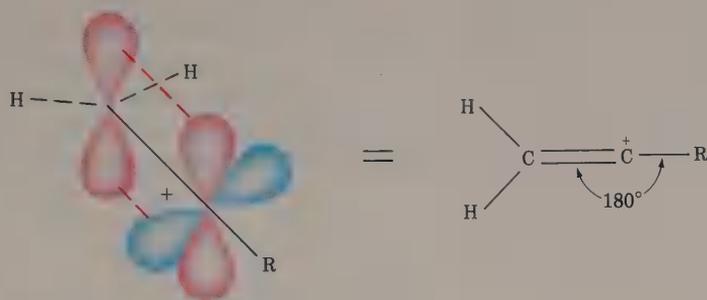
Vinyl chloride (chloroethylene), used as starting material for preparation of poly(vinyl chloride), was once produced on an immense industrial scale by mercuric chloride-catalyzed addition of HCl to acetylene. Now, however, other synthetic methods are used.



Bromine and chlorine also add to alkynes to give addition products, and trans stereochemistry again results:







**Figure 8.3** The electronic structure of a vinylic carbocation. The cationic carbon atom is *sp* hybridized.

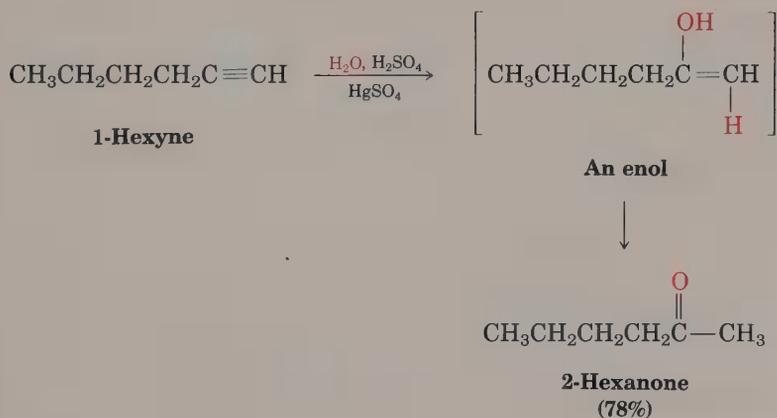
**PROBLEM** .....

**8.3** What products would you expect from these reactions?



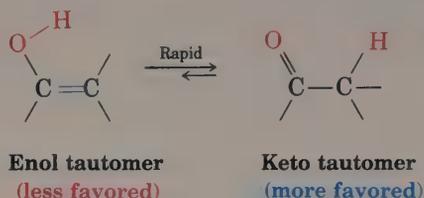
## 8.5 Hydration of Alkynes

Alkynes can't be hydrated as easily as alkenes because of their lower reactivity toward electrophilic addition. In the presence of mercuric sulfate catalyst, however, hydration occurs readily:

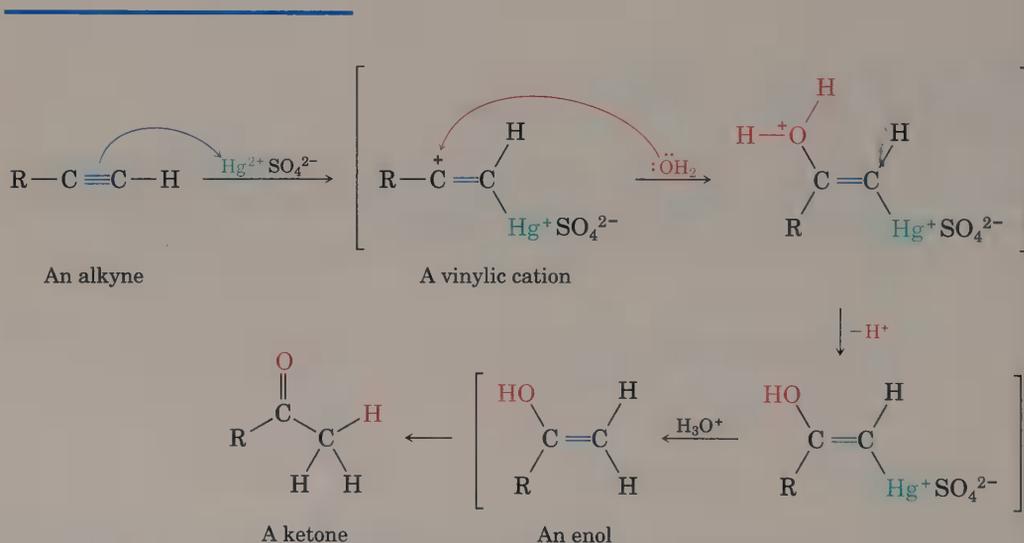


Markovnikov regiochemistry is found for the hydration with the OH group adding to the more highly substituted carbon and the H attaching to the less highly substituted carbon. Interestingly, though, the product is not

the expected vinylic alcohol, or **enol** (*ene* + *ol*), but is instead a ketone. Although the vinylic alcohol may well be an intermediate in the reaction, it immediately rearranges to a ketone in a process called **tautomerism**. **Tautomers** are special kinds of constitutional isomers that are readily interconvertible by rapid equilibration, and will be studied in more detail in Section 22.1. With few exceptions, the keto–enol tautomeric equilibrium lies heavily on the side of the ketone; enols are almost never isolated.

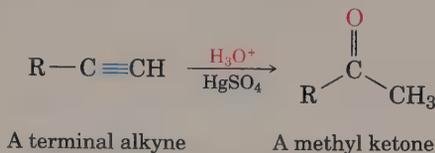
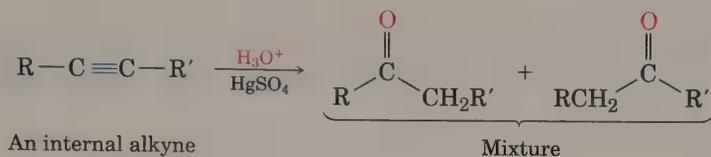


The mechanism of the mercuric ion-catalyzed alkyne hydration reaction is probably analogous to the oxymercuration reaction of alkenes (Section 7.4). Electrophilic addition of mercuric ion to the alkyne gives a vinylic cation, which reacts with water and loses a proton to yield an organomercury intermediate. In contrast to alkene oxymercuration, no treatment with  $\text{NaBH}_4$  is necessary to remove the mercury; the acidic reaction conditions alone are sufficient to allow replacement of mercury by hydrogen (Figure 8.4).



**Figure 8.4** Mechanism of the mercuric ion-catalyzed hydration of an alkyne to yield a ketone. The reaction yields an intermediate enol that rapidly tautomerizes to give a ketone.

A mixture of both possible ketones results when an internal alkyne is hydrated, and the reaction is therefore most useful when applied to terminal alkynes ( $\text{RC}\equiv\text{CH}$ ), since only methyl ketones are formed.

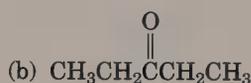
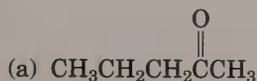


PROBLEM.....

8.4 What product would you obtain by hydration of 4-octyne? of 2-methyl-4-octyne?

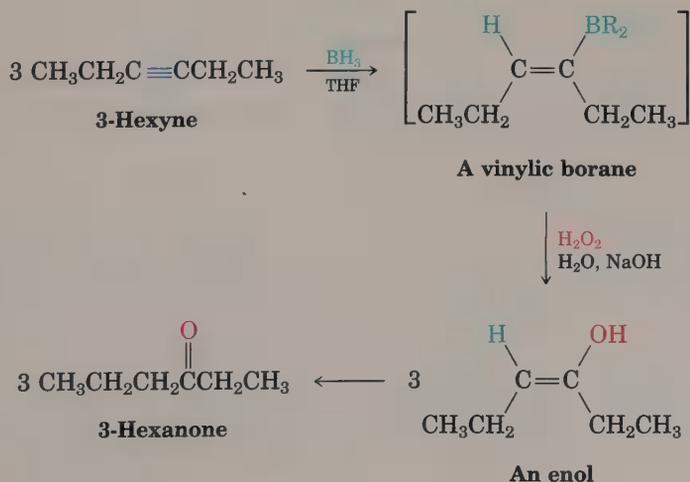
PROBLEM.....

8.5 What alkynes would you start with to prepare these ketones by a hydration reaction?

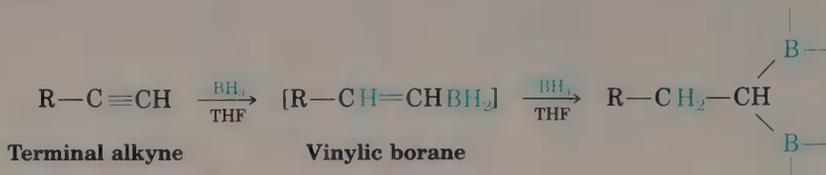


## 8.6 Hydroboration of Alkynes

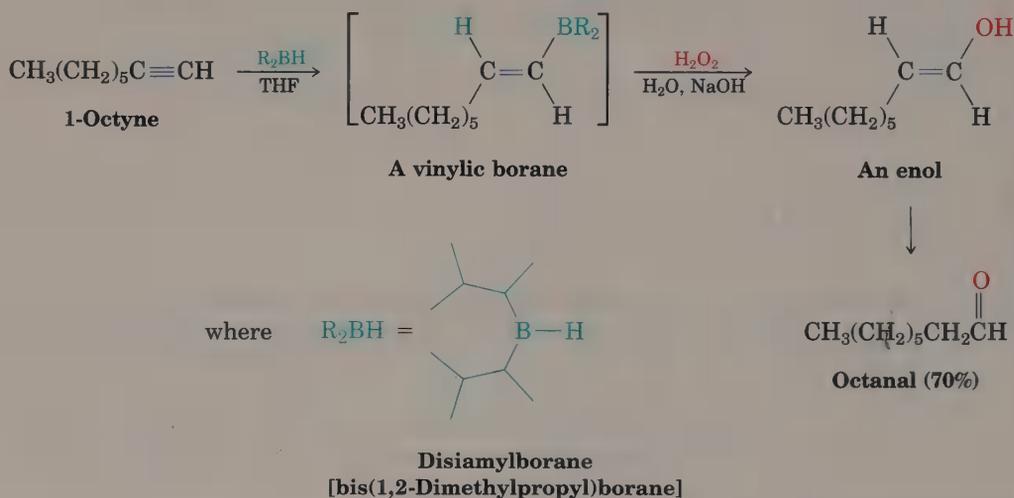
We saw in Section 7.5 that addition of  $\text{BH}_3$  to an alkene takes place in a single step through a four-membered-ring transition state without involving a carbocation intermediate. Borane also adds rapidly to alkynes, and the resulting vinylic boranes are quite useful. Hydroboration of symmetrically substituted internal alkynes gives vinylic boranes that are oxidized by basic hydrogen peroxide to yield ketones (via enols). Hydroboration of unsymmetrically substituted internal alkynes gives a mixture of two ketones:



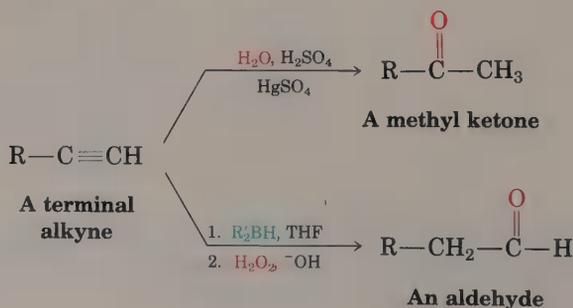
Terminal alkynes also react with borane, although in practice the reaction is difficult to stop at the vinylic borane stage. Under normal circumstances, a second addition of borane to the intermediate vinylic borane occurs:



To prevent this double addition, a bulky, sterically hindered borane such as bis(1,2-dimethylpropyl)borane (known commonly as disiamylborane) can be used in place of borane. When a terminal alkyne reacts with disiamylborane, addition of B-H to the C-C triple bond occurs with the expected non-Markovnikov regiochemistry. A second addition is rather slow, however, since the steric bulk of the large dialkylborane reagent makes approach to the double bond difficult. Oxidation of the vinylic borane intermediate then leads first to an enol and finally to an aldehyde.

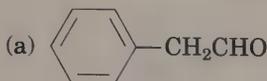


Note that the hydroboration-oxidation sequence is *complementary* to the direct hydration reaction of alkynes, since different products result. Direct hydration of a terminal alkyne with aqueous acid and mercuric sulfate leads to a methyl ketone, whereas hydroboration-oxidation of the same alkyne leads to an aldehyde:



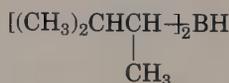
## PROBLEM

- 8.6 What alkynes would you start with to prepare these compounds by a hydroboration-oxidation reaction?



## PROBLEM

- 8.7 Disiamylborane is made by addition of  $\text{BH}_3$  to two equivalents of an alkene. What alkene is used?



Disiamylborane

## PROBLEM

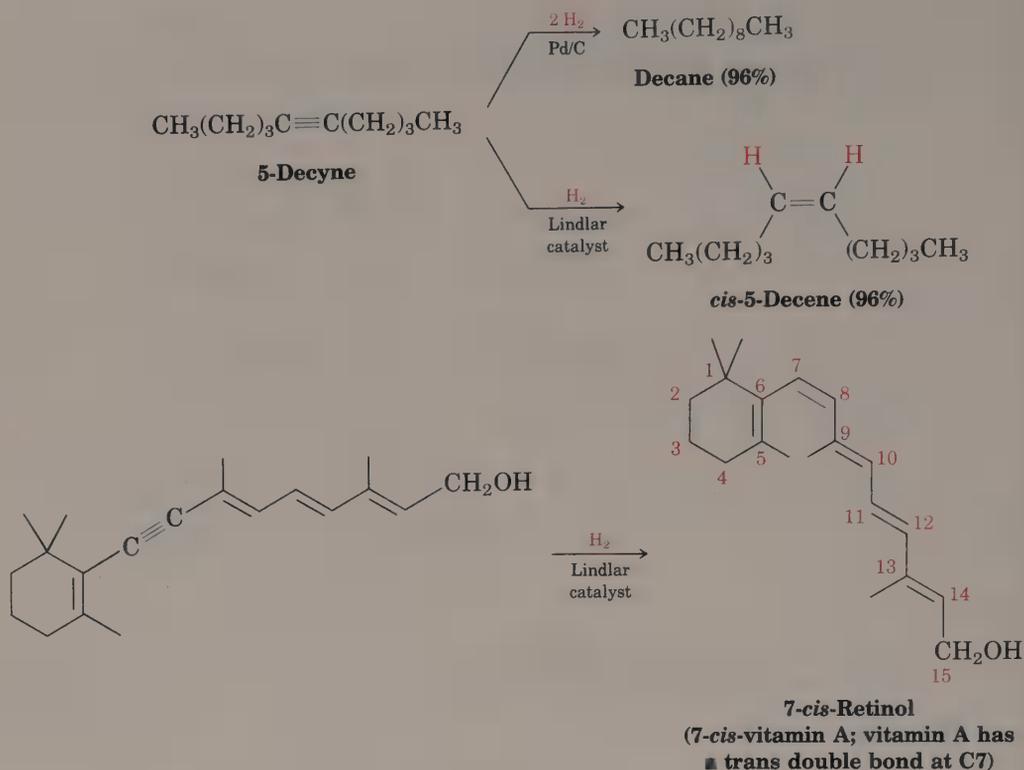
- 8.8 What organic product is formed in addition to the aldehyde when disiamylborane is used to hydroborate a terminal alkyne?

## 8.7 Reduction of Alkynes

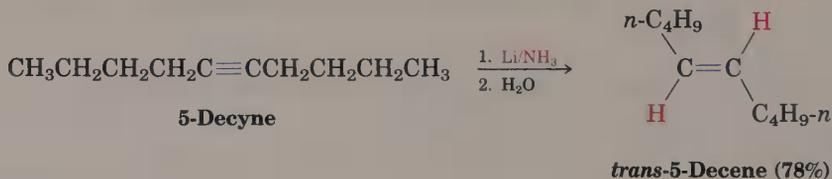
Alkynes are easily converted to alkanes by addition of hydrogen over a metal catalyst. Heat-of-hydrogenation data indicate that the first step in the reaction has a larger  $\Delta H^\circ$  than the second step, and we might therefore expect alkynes to reduce more readily than alkenes. This turns out to be the case.



Since alkynes reduce somewhat faster than alkenes, triple-bond hydrogenation can be stopped at the alkene stage if a suitable catalyst is used. The catalyst most often used for this purpose is the **Lindlar catalyst**, a finely divided palladium metal that has been precipitated onto a calcium carbonate support and then deactivated by treatment with lead acetate and quinoline, an aromatic amine. Since hydrogenation occurs with syn stereochemistry on the catalyst surface (Section 7.7), alkynes yield *cis* alkenes on reaction with one equivalent of hydrogen. This reaction has been explored extensively by the Hoffmann-LaRoche pharmaceutical company, where it is used in the commercial synthesis of vitamin A.



A second method for the conversion of alkynes to alkenes employs sodium or lithium metal in liquid ammonia as solvent. This method is complementary to the Lindlar reduction, since it produces *trans* rather than *cis* alkenes. Remarkably, alkali metals such as lithium and sodium dissolve in pure liquid ammonia at  $-33^\circ\text{C}$  to produce a deep blue solution. When an alkyne is added to this blue solution, reduction of the triple bond occurs and a *trans* alkene results. For example, 5-decyne gives *trans*-5-decene in good yield on treatment with lithium in liquid ammonia:



The mechanism of alkyne reduction by lithium in liquid ammonia involves donation of an electron to the triple bond to yield an intermediate **anion radical**—a species that is *both* a radical (has an odd number of electrons) and an anion (has a negative charge). This intermediate then takes a proton from ammonia to give a vinylic radical. Addition of a second electron to the vinylic radical gives a vinylic anion, which takes a second proton from ammonia to give *trans* alkene product (Figure 8.5).

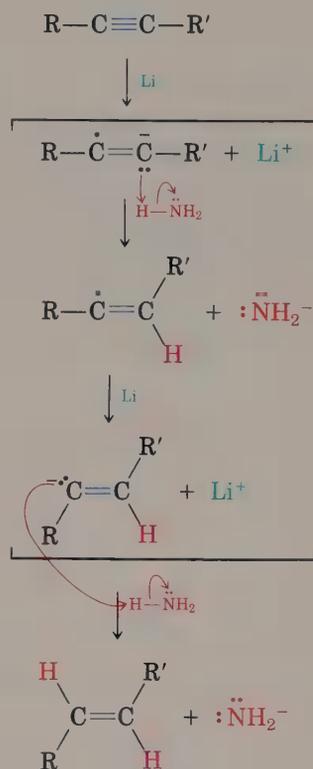
The *trans* stereochemistry of the final alkene product is established during the second reduction step when the less hindered, *trans* vinylic anion

Lithium metal donates an electron to the alkyne to give an anion radical . . .

. . . which abstracts a proton from ammonia solvent to yield a vinylic radical.

The vinylic radical accepts another electron from a second lithium atom to produce a vinylic anion . . .

. . . which abstracts another proton from ammonia solvent to yield the final *trans* alkene product.



**Figure 8.5** Mechanism of the lithium/ammonia reduction of an alkyne to produce a *trans* alkene.

is formed from the vinylic radical. Vinylic radicals undergo rapid *cis*–*trans* equilibration, but vinylic anions equilibrate much less rapidly. Thus, the more stable *trans* (rather than less stable *cis*) vinylic anion is formed and then protonated before it can equilibrate.

PROBLEM . . . . .

- 8.9 Starting from any alkyne you choose, how would you prepare these alkenes?  
 (a) *trans*-2-Octene                      (b) *cis*-3-Heptene                      (c) 3-Methyl-1-pentene
- . . . . .

## 8.8 Alkyne Acidity: Formation of Acetylide Anions

The most striking difference between the chemistry of alkenes and alkynes is that *terminal alkynes are weakly acidic*. When a terminal alkyne is treated with a strong base such as sodium amide,  $NaNH_2$ , the terminal hydrogen is removed, and an **acetylide anion** is formed:

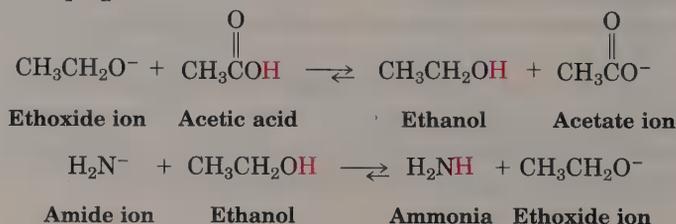


According to the Brønsted–Lowry definition, an acid is a species that donates a proton. Although we usually think of oxyacids ( $\text{H}_2\text{SO}_4$ ,  $\text{H}_2\text{O}$ ) or halogen acids ( $\text{HCl}$ ,  $\text{HBr}$ ) in this context, *any* compound containing a hydrogen atom can be considered an acid under the proper circumstances. As discussed earlier (Section 2.6), we can establish an acidity order by measuring dissociation constants of acids and expressing the results as  $\text{p}K_{\text{a}}$  values (Table 8.1). Recall from Section 2.6 that a low  $\text{p}K_{\text{a}}$  corresponds to a strong acid, and a high  $\text{p}K_{\text{a}}$  corresponds to a weak acid.

**TABLE 8.1** Strengths of Some Common Acids ( $\text{p}K_{\text{a}}$  Values)

Acid	Formula	$\text{p}K_{\text{a}}$	
Sulfuric acid	$\text{HO}-\overset{\text{O}}{\parallel}{\text{S}}-\text{OH}$	$\sim -9$	Stronger acid
Acetic acid	$\text{CH}_3\overset{\text{O}}{\parallel}{\text{C}}\text{OH}$	4.72	
Water	$\text{HOH}$	15.74	
Ethanol	$\text{CH}_3\text{CH}_2\text{OH}$	16	
Acetylene	$\text{HC}\equiv\text{CH}$	25	
Ammonia	$\text{H}_2\text{NH}$	35	

At one end of the scale, the mineral acids such as sulfuric acid are very strong. Carboxylic acids such as acetic acid are moderately strong, water and alcohols such as ethanol are intermediate, and ammonia is a very weak acid. Since a stronger acid donates its proton to the anion of a weaker acid in an acid–base reaction, a rank-ordered list tells us what bases are needed to deprotonate what acids. For example, since acetic acid ( $\text{p}K_{\text{a}} = 4.72$ ) is a stronger acid than ethanol ( $\text{p}K_{\text{a}} = 16$ ), we know that the anion of ethanol (ethoxide ion,  $\text{CH}_3\text{CH}_2\text{O}^-$ ) will remove a proton from acetic acid. Similarly, amide ion ( $^-\text{NH}_2$ ), the anion of ammonia ( $\text{p}K_{\text{a}} = 35$ ), will remove a proton from ethanol ( $\text{p}K_{\text{a}} = 16$ ).



In general:



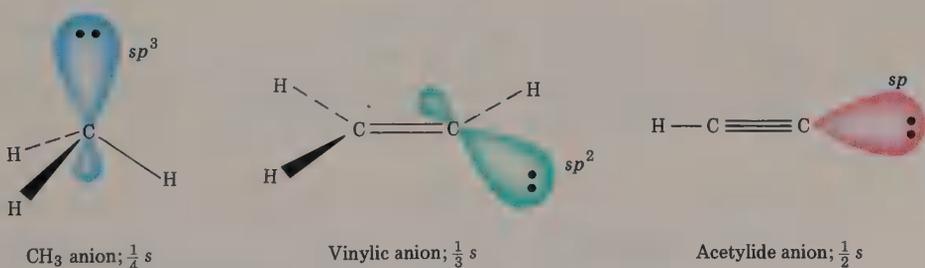
Where do hydrocarbons lie on the acidity scale? As the data in Table 8.2 indicate, both methane ( $pK_a \approx 60$ ) and ethylene ( $pK_a = 44$ ) are very weak acids. For all practical purposes, neither can be deprotonated by a base. Acetylene, however, has a  $pK_a$  of 25 and is thus easily deprotonated by a sufficiently strong base. The anion of any acid whose  $pK_a$  is greater than 25 will abstract a proton from acetylene. Amide ion,  $\text{NH}_2^-$ , for example, is the conjugate base of ammonia ( $\text{NH}_3$ ;  $pK_a = 35$ ) and is therefore able to abstract a proton from terminal alkynes.

TABLE 8.2 Acidity of Simple Hydrocarbons

Type	Example	$K_a$	$pK_a$
Alkyne	$\text{HC}\equiv\text{CH}$	$10^{-25}$	25
Alkene	$\text{H}_2\text{C}=\text{CH}_2$	$10^{-44}$	44
Alkane	$\text{CH}_4$	$\approx 10^{-60}$	$\approx 60$

↑ Stronger acid  
↓ Weaker acid

What makes terminal alkynes so acidic? The best intuitive explanation is simply that acetylide anions have more  $s$  character in the orbital occupied by the electron pair than do alkyl or vinylic anions. Alkane anions are  $sp^3$  hybridized, so the negative charge resides in an orbital that has one-quarter “ $s$  character” and three-quarters  $p$  character; vinylic anions are  $sp^2$  hybridized and therefore have one-third  $s$  character; and acetylide anions ( $sp$ ) have one-half  $s$  character. Since  $s$  orbitals are lower in energy and nearer the positively charged nucleus than  $p$  orbitals, a negative charge is stabilized to a greater extent in an orbital with high  $s$  character than in an orbital with low  $s$  character (Figure 8.6).



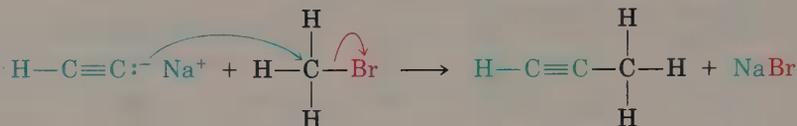
**Figure 8.6** A comparison of alkyl, vinylic, and acetylide anions. The acetylide anion, with  $sp$  hybridization, has more  $s$  character and is more stable.

## PROBLEM .....

- 8.10 The  $pK_a$  of acetone,  $\text{CH}_3\text{COCH}_3$ , is 20. Which of the following bases are strong enough to deprotonate acetone?
- (a)  $\text{KOH}$  ( $pK_a$  of  $\text{H}_2\text{O} = 15.7$ )      (b)  $\text{Na}^+ \text{ } ^-\text{C}\equiv\text{CH}$  ( $pK_a$  of  $\text{C}_2\text{H}_2 = 25$ )  
 (c)  $\text{NaHCO}_3$  ( $pK_a$  of  $\text{H}_2\text{CO}_3 = 6.4$ )      (d)  $\text{Na}^+ \text{ } ^-\text{OCH}_3$  ( $pK_a$  of  $\text{CH}_3\text{OH} = 15.6$ )
- .....

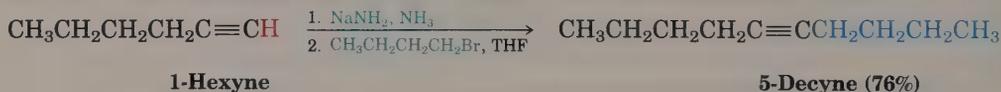
## 8.9 Alkylation of Acetylide Anions

The presence of an unshared electron pair on carbon makes acetylide anions strongly nucleophilic. As a result, acetylide anions in tetrahydrofuran solvent react with alkyl halides such as bromomethane to substitute for the halogen and yield a new alkyne product:

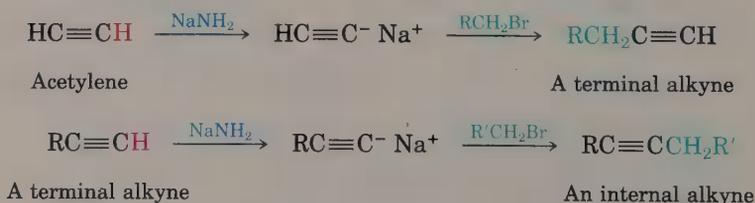


Although we won't study the details of this substitution reaction until Chapter 11, we can picture it as happening by the pathway shown in Figure 8.7. The nucleophilic acetylide ion attacks the positively polarized (and therefore electrophilic) carbon atom of bromomethane and pushes out bromide ion with the electron pair from the former C–Br bond, yielding propyne as product. We call such a reaction an **alkylation**, since a new alkyl group has become attached to the starting alkyne.

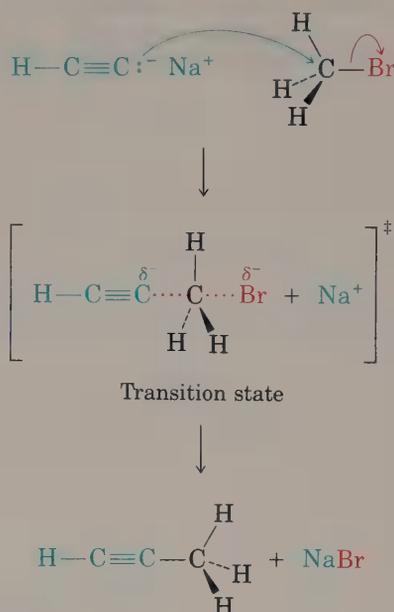
Alkyne alkylation is not limited to acetylide ion. *Any* terminal alkyne can be converted into its corresponding acetylide anion and alkylated by treatment with an alkyl halide to yield an internal alkyne product. For example, conversion of 1-hexyne into its anion, followed by reaction with 1-bromobutane, yields 5-decyne:



Because of its generality, acetylide alkylation is the best method for preparing substituted alkynes from simpler precursors. Terminal alkynes can be prepared by alkylation of acetylene itself, and internal alkynes can be prepared by further alkylation of a terminal alkyne.



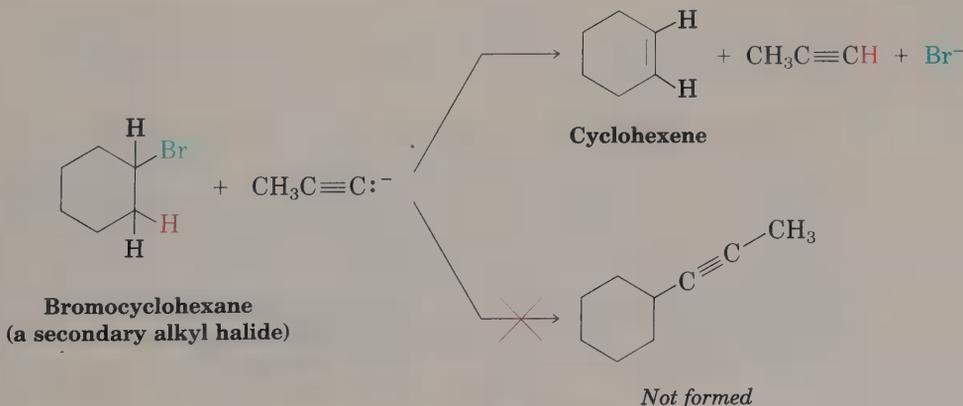
The nucleophilic acetylide anion uses its electron lone pair to form a bond to the positively polarized, electrophilic carbon atom of bromomethane. As the new C-C bond begins to form, the C-Br bond begins to break in the transition state.



The new C-C bond is fully formed and the old C-Br bond is fully broken at the end of the reaction.

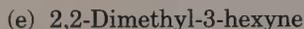
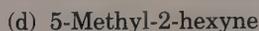
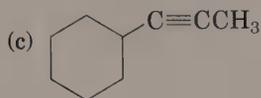
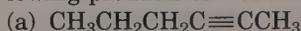
**Figure 8.7** A possible mechanism for the substitution reaction of acetylide anion with bromomethane.

Acetylide ion alkylation is limited to the use of primary alkyl bromides and iodides,  $\text{RCH}_2\text{X}$ , for reasons that will be discussed in more detail in Chapter 11. In addition to their reactivity as nucleophiles, acetylide ions are sufficiently strong bases to cause dehydrohalogenation instead of substitution when they react with secondary and tertiary alkyl halides. For example, reaction of bromocyclohexane with propyne anion leads to formation of the elimination product cyclohexene rather than the substitution product cyclohexylpropyne.



## PROBLEM.....

8.11 Show the terminal alkyne and alkyl halide starting materials from which the following products can be obtained. Where two routes look feasible, list both choices.

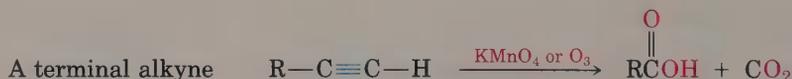
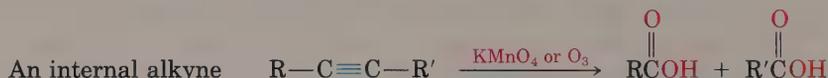


## PROBLEM.....

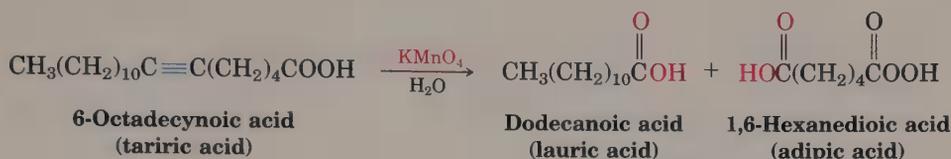
8.12 How would you prepare *cis*-2-butene, starting from 1-propyne, an alkyl halide, and any other reagents needed? (This problem can't be worked in a single step. You'll have to carry out more than one chemical reaction.)

## 8.10 Oxidative Cleavage of Alkynes

Alkynes, like alkenes, can be cleaved by reaction with powerful oxidizing agents such as potassium permanganate or ozone. A triple bond is generally less reactive than a double bond, however, and yields of cleavage products are sometimes low. The reaction is too complex mechanistically to discuss in detail, but the products obtained are carboxylic acids. If a terminal alkyne is oxidized,  $\text{CO}_2$  is formed as one product.

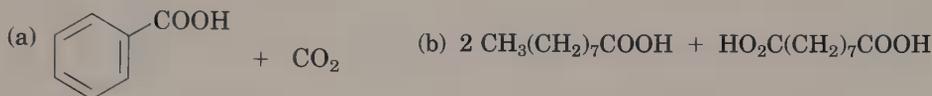


The major application of alkyne oxidation reactions is in structure determination of substances isolated from natural sources. For example, when the compound called tariric acid was isolated from the Guatemalan plant *Picramnia tariri*, it was identified as an 18-carbon straight-chain acetylenic acid, but the position of the triple bond in the chain was unknown. Since oxidation of tariric acid with potassium permanganate gave two products identified as dodecanoic acid and hexanedioic acid, the position of the triple bond could be established.



PROBLEM.....

- 8.13 Propose structures for alkynes that give the following products on oxidative cleavage by  $\text{KMnO}_4$ :



## 8.11 Organic Synthesis

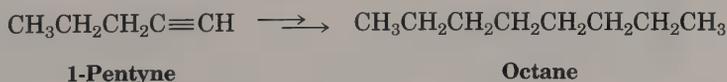
The laboratory synthesis of an organic molecule from simple precursors might be carried out for many reasons. In the pharmaceutical industry, new organic molecules are designed and synthesized in the hope that some might be useful new drugs. In the chemical industry, synthesis is often undertaken to devise more economical routes to known compounds. In academic laboratories, the synthesis of highly complex molecules is sometimes done purely for the intellectual challenge involved in mastering so difficult a subject. The successful synthesis route is a highly creative work that is sometimes described by such subjective terms as *elegant* or *beautiful*.

In this book, too, we will often devise syntheses of molecules from simpler precursors. The purpose, however, is purely pedagogical. The ability to plan workable synthetic sequences demands a thorough knowledge of a wide variety of organic reactions. Furthermore, it requires more than a theoretical knowledge of reactions; it also requires a practical grasp for the proper fitting together of steps in a sequence such that each reaction does only what is desired. *Working synthesis problems is an excellent way to learn organic chemistry.*

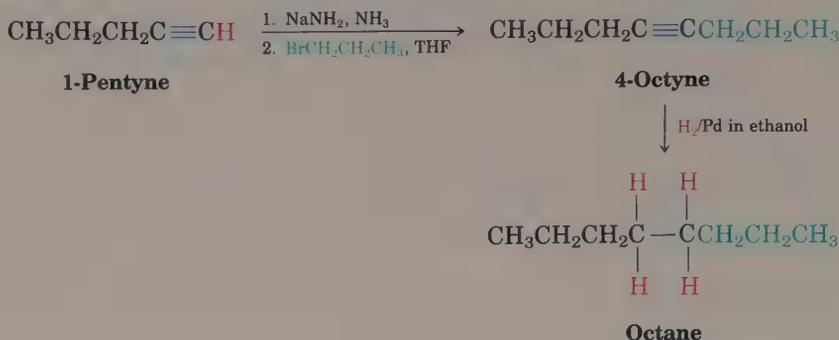
Some of the syntheses we plan may appear trivial. Here's an example:

PRACTICE PROBLEM.....

Prepare octane from 1-pentyne.



**Solution** First alkylate the acetylide anion of 1-pentyne with 1-bromopropane, and then reduce the product using catalytic hydrogenation:



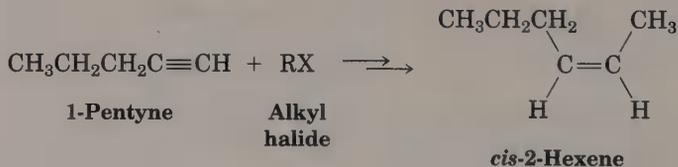
Although the synthesis route just presented should work perfectly well, it has little practical value since a chemist can simply *buy* octane from several dozen chemical supply companies. The value of working the problem is that it makes us approach a chemical problem in a logical way, draw on our knowledge of chemical reactions, and organize that knowledge into a workable plan—it helps us *learn* organic chemistry.

There's no secret to planning an organic synthesis. All it takes is a knowledge of the different reactions and a lot of practice. But here's a hint: *Work backward*. Look at the final product and ask, "What was the *immediate* precursor of that product?" For example, if the end product is an alkyl halide, the immediate precursor might be an alkene (via HX addition). Having found an immediate precursor, proceed backward again, one step at a time, until a suitable starting material is found.

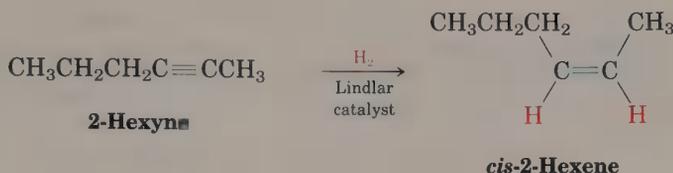
Let's work some examples of increasing complexity.

#### PRACTICE PROBLEM.....

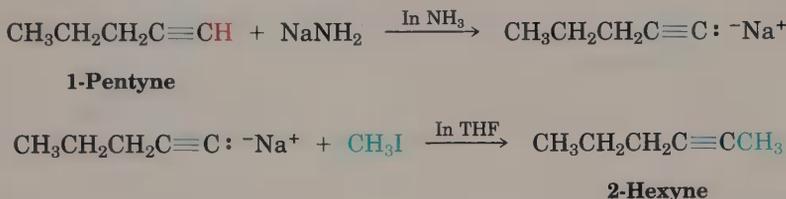
Starting from 1-pentyne and any alkyl halide needed, synthesize *cis*-2-hexene. More than one step is required.



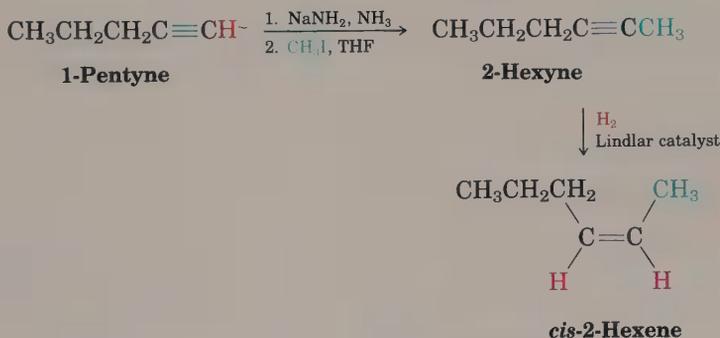
**Solution** First ask "What is an immediate precursor of a *cis*-disubstituted alkene?" We know that alkenes can be prepared from alkynes by partial reduction. The proper choice of experimental conditions will allow us to prepare either a *trans*-disubstituted alkene (using lithium in liquid ammonia) or a *cis*-disubstituted alkene (using catalytic hydrogenation over the Lindlar catalyst). Thus, reduction of 2-hexyne by catalytic hydrogenation using the Lindlar catalyst should yield *cis*-2-hexene:



Next ask "What is an immediate precursor of 2-hexyne?" We've seen that internal alkynes can be prepared by alkylation of terminal alkyne anions (acetylides). In the present instance, we're told to start with 1-pentyne. Thus, alkylation of the anion of 1-pentyne with iodomethane should yield 2-hexyne:

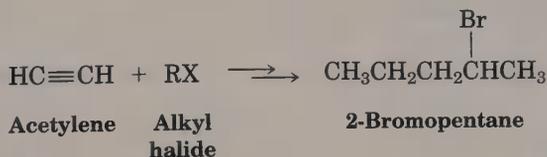


In three steps we've synthesized *cis*-2-hexene from the given starting materials:

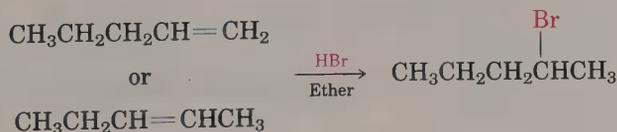


### PRACTICE PROBLEM.....

Starting from acetylene and any alkyl halide needed, synthesize 2-bromopentane. More than one step is required.

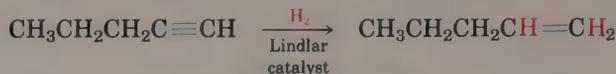


**Solution** First ask "What is an immediate precursor of an alkyl halide?" Perhaps an alkene:



Of the two alkene possibilities, addition of HBr to 1-pentene looks like a better choice than addition to 2-pentene, since the latter would give a mixture of isomers.

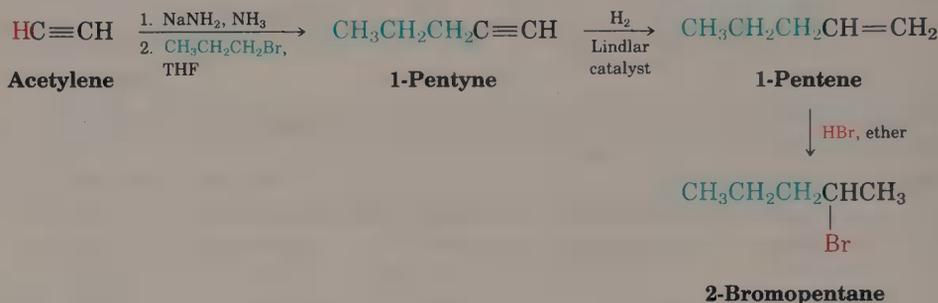
Next ask "What is an immediate precursor of an alkene?" Perhaps an alkyne, which could be partially reduced:



Next ask "What is an immediate precursor of a terminal alkyne?" Perhaps sodium acetylide and an alkyl halide:

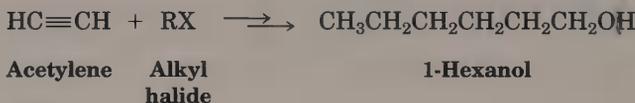


In four steps we have synthesized the desired material from acetylene and 1-bromopropane.

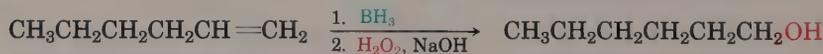


### PRACTICE PROBLEM.....

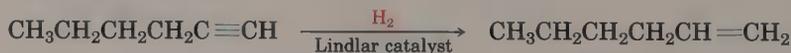
Synthesize 1-hexanol from acetylene and any alkyl halide needed.



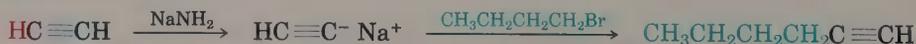
**Solution** "What is an immediate precursor of a primary alcohol?" Perhaps an alkene, which could be hydrated by reaction with borane followed by oxidation with  $\text{H}_2\text{O}_2$ :



"What is an immediate precursor of a terminal alkene?" Perhaps a terminal alkyne, which could be reduced:



"What is an immediate precursor of 1-hexyne?" Perhaps acetylene and 1-bromobutane:



We have completed the synthesis in three steps by working backward.

## PROBLEM.....

- 8.14 Beginning with 4-octyne as your only source of carbon, and using any inorganic reagents necessary, how would you synthesize the following compounds?
- (a) Butanoic acid                      (b) *cis*-4-Octene                      (c) 4-Bromooctane  
 (d) 4-Octanol (4-hydroxyoctane)    (e) 4,5-Dichlorooctane

## PROBLEM.....

- 8.15 Beginning with acetylene and any alkyl halides needed, how would you synthesize the following compounds?
- (a) Decane    (b) 2,2-Dimethylhexane    (c) Hexanal    (d) 2-Heptanone

## 8.12 Summary and Key Words

**Alkynes** are hydrocarbons that contain one or more carbon–carbon triple bonds. Alkyne carbon atoms are *sp* hybridized, and the triple bond is formed by one *sp-sp* sigma bond and two *p-p* pi bonds. There are relatively few general methods of alkyne synthesis. The two best are the alkylation of acetylide anions with a primary alkyl halide and the twofold elimination of HX from vicinal dihalides.

The chemistry of alkynes is dominated by electrophilic addition reactions similar to those of alkenes. For example, alkynes react with HBr and HCl to yield **vinylic** halides, and with Br<sub>2</sub> and Cl<sub>2</sub> to yield 1,2-dihalides (**vicinal** dihalides). Alkynes can be hydrated (addition of H<sub>2</sub>O) by either of two procedures. Reaction of an alkyne with aqueous sulfuric acid in the presence of mercuric ion catalyst leads to an intermediate **enol** that immediately isomerizes (**tautomerizes**) to yield a ketone. Since the addition reaction occurs with Markovnikov regiochemistry, a methyl ketone is produced from a terminal alkyne. Alternatively, addition of water can be effected by hydroboration of an alkyne, followed by oxidation with basic hydrogen peroxide. Disiamylborane, a sterically hindered dialkylborane, is often used in this reaction, allowing aldehydes to be prepared in good yield from terminal alkynes.

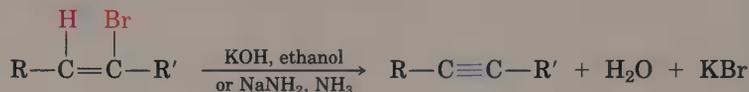
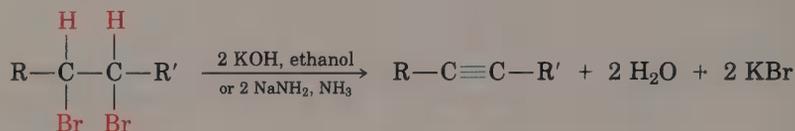
Alkynes can also be reduced to yield alkenes and alkanes. Complete reduction of the triple bond over a normal palladium hydrogenation catalyst yields an alkane, whereas partial reduction by catalytic hydrogenation over a **Lindlar catalyst** yields a *cis* alkene. Reduction of the alkyne with lithium in ammonia yields a *trans* alkene.

One of the most striking differences in chemical reactivity between alkynes and alkenes is due to their different acidities. Terminal alkynes contain an acidic hydrogen that can be removed by a strong base to yield an **acetylide anion**. Acetylide anions act as nucleophiles and can displace halide ion from primary alkyl halides in an **alkylation** reaction. Acetylide anions are more stable than either alkyl anions or vinylic anions because their negative charge is in a hybrid orbital with much *s* character, allowing the charge to be closer to the nucleus.

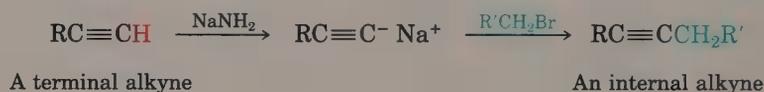
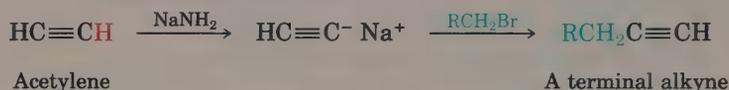
## 8.13 Summary of Reactions

### 1. Preparation of alkynes

#### a. Dehydrohalogenation of vicinal dihalides (Section 8.3)

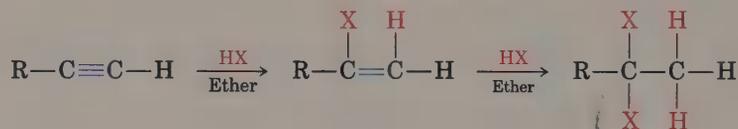


#### b. Acetylide ion alkylation (Section 8.9)

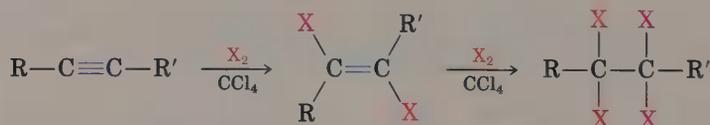


### 2. Reactions of alkynes

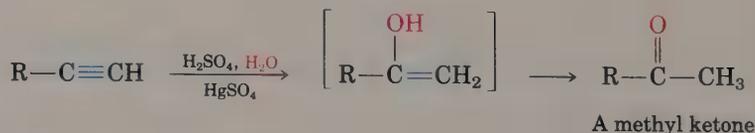
#### a. Addition of HX, where X = Br or Cl (Section 8.4)



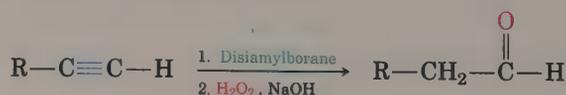
#### b. Addition of X<sub>2</sub>, where X = Br or Cl (Section 8.4)



#### c. Mercuric sulfate-catalyzed hydration (Section 8.5)

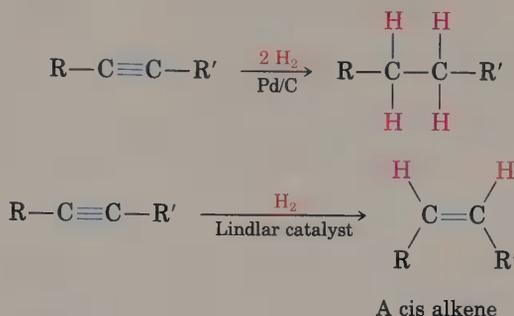


#### d. Hydroboration-oxidation (Section 8.6)

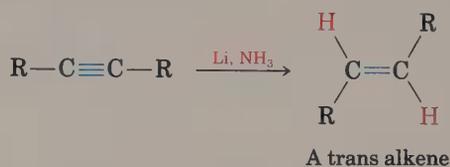


## e. Reduction (Section 8.7)

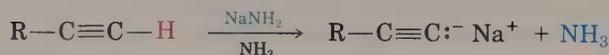
## (1) Catalytic hydrogenation



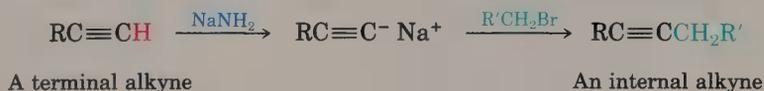
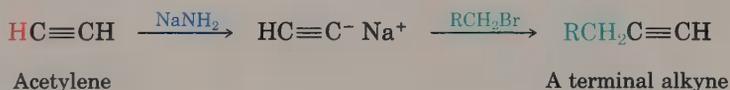
## (2) Lithium/ammonia



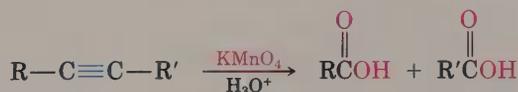
## f. Acidity: conversion into acetylide anions (Section 8.8)



## g. Acetylide ion alkylation (Section 8.9)

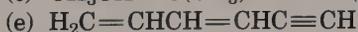
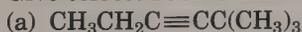


## h. Oxidative cleavage (Section 8.10)

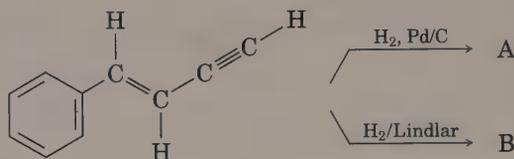


## ADDITIONAL PROBLEMS .....

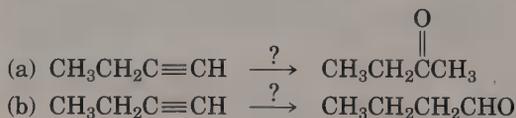
8.16 Give correct IUPAC names for these compounds:

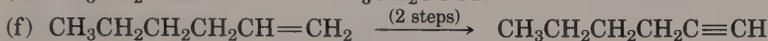
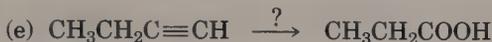
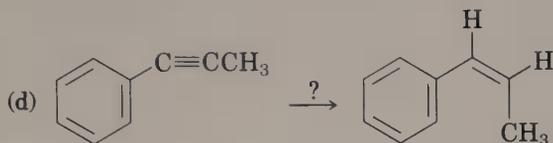
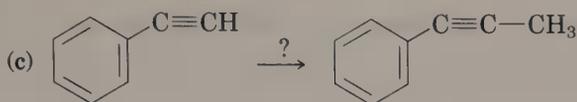


- 8.17 Draw structures corresponding to these names:
- (a) 3,3-Dimethyl-4-octyne (b) 3-Ethyl-5-methyl-1,6,8-decatriyne  
 (c) 2,2,5,5-Tetramethyl-3-hexyne (d) 3,4-Dimethylcyclodecyne  
 (e) 3,5-Heptadien-1-yne (f) 3-Chloro-4,4-dimethyl-1-nonen-6-yne  
 (g) 3-*sec*-Butyl-1-heptyne (h) 5-*tert*-Butyl-2-methyl-3-octyne
- 8.18 The following names are incorrect. Draw the structures and give the correct names.
- (a) 1-Ethyl-5,5-dimethyl-1-hexyne (b) 2,5,5-Trimethyl-6-heptyne  
 (c) 3-Methylhept-5-en-1-yne (d) 2-Isopropyl-5-methyl-7-octyne  
 (e) 3-Hexen-5-yne (f) 5-Ethynyl-1-methylcyclohexane
- 8.19 These two hydrocarbons have been isolated from various plants in the sunflower family. Name them according to IUPAC rules.
- (a)  $\text{CH}_3\text{CH}=\text{CHC}\equiv\text{CC}\equiv\text{CCH}=\text{CHCH}=\text{CHCH}=\text{CH}_2$  (all trans)  
 (b)  $\text{CH}_3\text{C}\equiv\text{CC}\equiv\text{CC}\equiv\text{CC}\equiv\text{CC}\equiv\text{CCH}=\text{CH}_2$
- 8.20 Predict the products of these reactions:

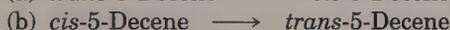
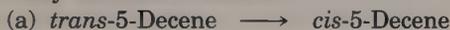


- 8.21 A hydrocarbon of unknown structure has the formula  $\text{C}_8\text{H}_{10}$ . On catalytic hydrogenation over the Lindlar catalyst, one equivalent of  $\text{H}_2$  is absorbed. On hydrogenation over a palladium catalyst, however, three equivalents of  $\text{H}_2$  are absorbed.
- (a) How many rings/double bonds/triple bonds are present in the unknown?  
 (b) How many triple bonds are present?  
 (c) How many double bonds are present?  
 (d) How many rings are present?
- Explain your answers, and draw a possible structure that fits the data.
- 8.22 Predict the products from reaction of 1-hexyne with these reagents:
- (a) One equivalent HBr (b) One equivalent  $\text{Cl}_2$   
 (c)  $\text{H}_2$ , Lindlar catalyst (d)  $\text{NaNH}_2$  in  $\text{NH}_3$ , then  $\text{CH}_3\text{Br}$   
 (e)  $\text{H}_2\text{O}$ ,  $\text{H}_2\text{SO}_4$ ,  $\text{HgSO}_4$  (f) Two equivalents HCl
- 8.23 Predict the products from reaction of 5-decyne with these reagents:
- (a)  $\text{H}_2$ , Lindlar catalyst (b) Li in  $\text{NH}_3$   
 (c) One equivalent  $\text{Br}_2$  (d)  $\text{BH}_3$  in THF, then  $\text{H}_2\text{O}_2$   
 (e)  $\text{H}_2\text{O}$ ,  $\text{H}_2\text{SO}_4$ ,  $\text{HgSO}_4$  (f) Excess  $\text{H}_2$ , Pd/C catalyst
- 8.24 Predict the products from reaction of 2-hexyne with these reagents:
- (a) Two equivalents  $\text{Br}_2$  (b) One equivalent HBr  
 (c) Excess HBr (d) Li in  $\text{NH}_3$   
 (e)  $\text{H}_2\text{O}$ ,  $\text{H}_2\text{SO}_4$ ,  $\text{HgSO}_4$
- 8.25 Acetonitrile,  $\text{CH}_3\text{CN}$ , contains a carbon–nitrogen triple bond. Sketch the orbitals involved in the bonding in acetonitrile, and indicate the hybridization of each atom.
- 8.26 How would you carry out these reactions?

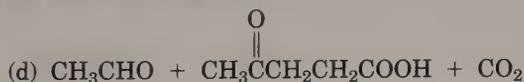
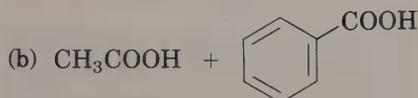




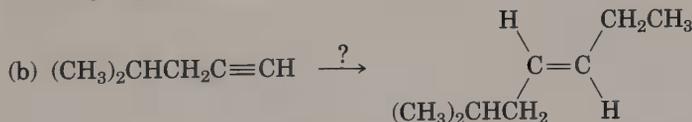
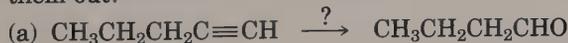
**8.27** Occasionally, chemists need to invert the stereochemistry of an alkene. That is, one might want to convert a *cis* alkene to a *trans* alkene, or vice versa. Although there is no one-step method for doing this alkene inversion, the transformation can be carried out by combining several reactions in the proper sequence. How would you carry out these reactions?



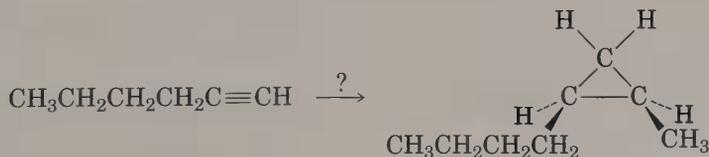
**8.28** Propose structures for hydrocarbons that give the following products on oxidative cleavage by  $\text{KMnO}_4$  or  $\text{O}_3$ .



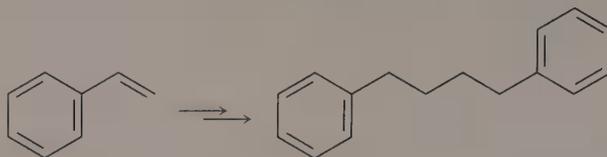
**8.29** Each of the following syntheses requires more than one step. How would you carry them out?



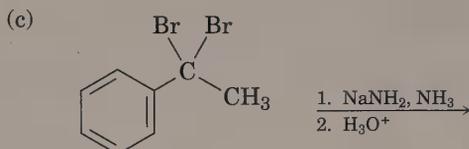
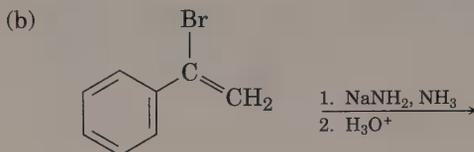
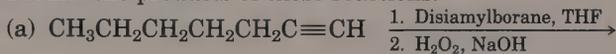
**8.30** How would you carry out the following transformation? More than one step is required.



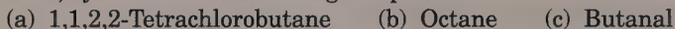
8.31 How would you carry out the following conversion? More than one step is needed.



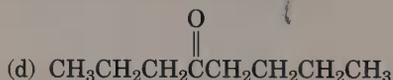
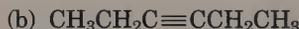
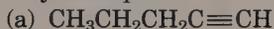
8.32 Predict the products of these reactions:



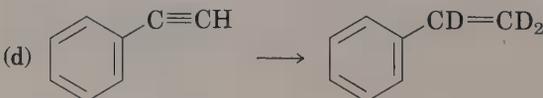
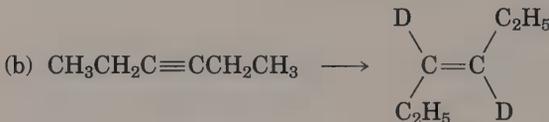
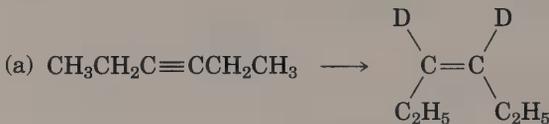
8.33 Using 1-butyne as the only source of carbon, along with any inorganic reagents you need, synthesize the following compounds. More than one step may be needed.



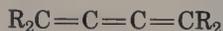
8.34 Using acetylene and any alkyl halides that have four or fewer carbons as starting materials, how would you synthesize the following compounds? More than one step may be required.



8.35 How would you carry out these reactions to introduce deuterium into organic molecules?

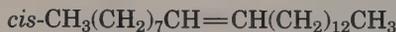


- 8.36 A *cumulene* is a compound with three adjacent double bonds. Draw an orbital picture of a cumulene. What kind of hybridization do the two central carbon atoms have? What is the geometric relationship of the substituents on one end to the substituents on the other end? What kind of isomerism is possible? Make a model to help you see the answer.



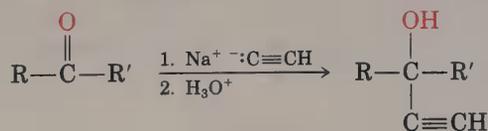
A cumulene

- 8.37 Although it is geometrically impossible for a triple bond to exist in a small ring, large-ring cycloalkynes are quite stable. How would you prepare cyclodecyne starting from acetylene and any alkyl halide needed?
- 8.38 The sex attractant given off by the common housefly is a simple alkene named *muscalure*. Propose a synthesis of muscalure starting from acetylene and any alkyl halides. What is the IUPAC name for muscalure?



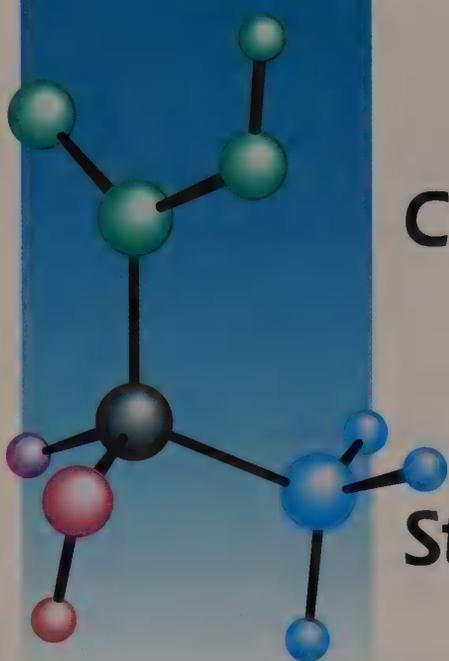
Muscalure

- 8.39 Compound A ( $C_9H_{12}$ ) absorbed three equivalents of hydrogen on catalytic reduction over a palladium catalyst to give B ( $C_9H_{18}$ ). On ozonolysis, compound A gave, among other things, a ketone that was identified as cyclohexanone. On treatment with  $NaNH_2$  in  $NH_3$ , followed by addition of iodomethane, compound A gave a new hydrocarbon, C ( $C_{10}H_{14}$ ). What are the structures of A, B, and C?
- 8.40 Hydrocarbon A has the formula  $C_{12}H_8$ . It absorbs eight equivalents of hydrogen on catalytic reduction over a palladium catalyst. On ozonolysis only two products are formed: oxalic acid ( $HOOC-COOH$ ) and succinic acid ( $HOOCCH_2CH_2COOH$ ). Formulate these reactions and propose a structure for A.
- 8.41 Organometallic reagents such as sodium acetylide undergo an addition reaction with ketones, giving alcohols:



How might you use this reaction to prepare 2-methyl-1,3-butadiene, the starting material used in the manufacture of synthetic rubber?

- 8.42 Erythrogenic acid,  $C_{18}H_{26}O_2$ , is an interesting acetylenic fatty acid that turns a vivid red on exposure to light. On catalytic hydrogenation over a palladium catalyst, five equivalents of hydrogen are absorbed, and stearic acid,  $CH_3(CH_2)_{16}COOH$ , is produced. Ozonolysis of erythrogenic acid gives four products: formaldehyde,  $CH_2O$ ; oxalic acid,  $HOOC-COOH$ ; azelaic acid,  $HOOC(CH_2)_7COOH$ ; and the aldehyde acid  $OHC(CH_2)_4COOH$ . With this information, draw two possible structures for erythrogenic acid. Suggest a way to tell them apart by carrying out some simple reactions.



## Chapter 9

# Stereochemistry

Are you right-handed or left-handed? Though most of us don't often think about it, handedness plays a surprisingly large role in our daily activities. Many musical instruments, such as oboes and clarinets, have a handedness to them; the last available softball glove always fits the wrong hand; left-handed people write in a "funny" way. The fundamental reason for these difficulties, of course, is that our hands aren't identical—they're *mirror images*. When you hold a *right* hand up to a mirror, the image you see looks like a *left* hand. Try it.

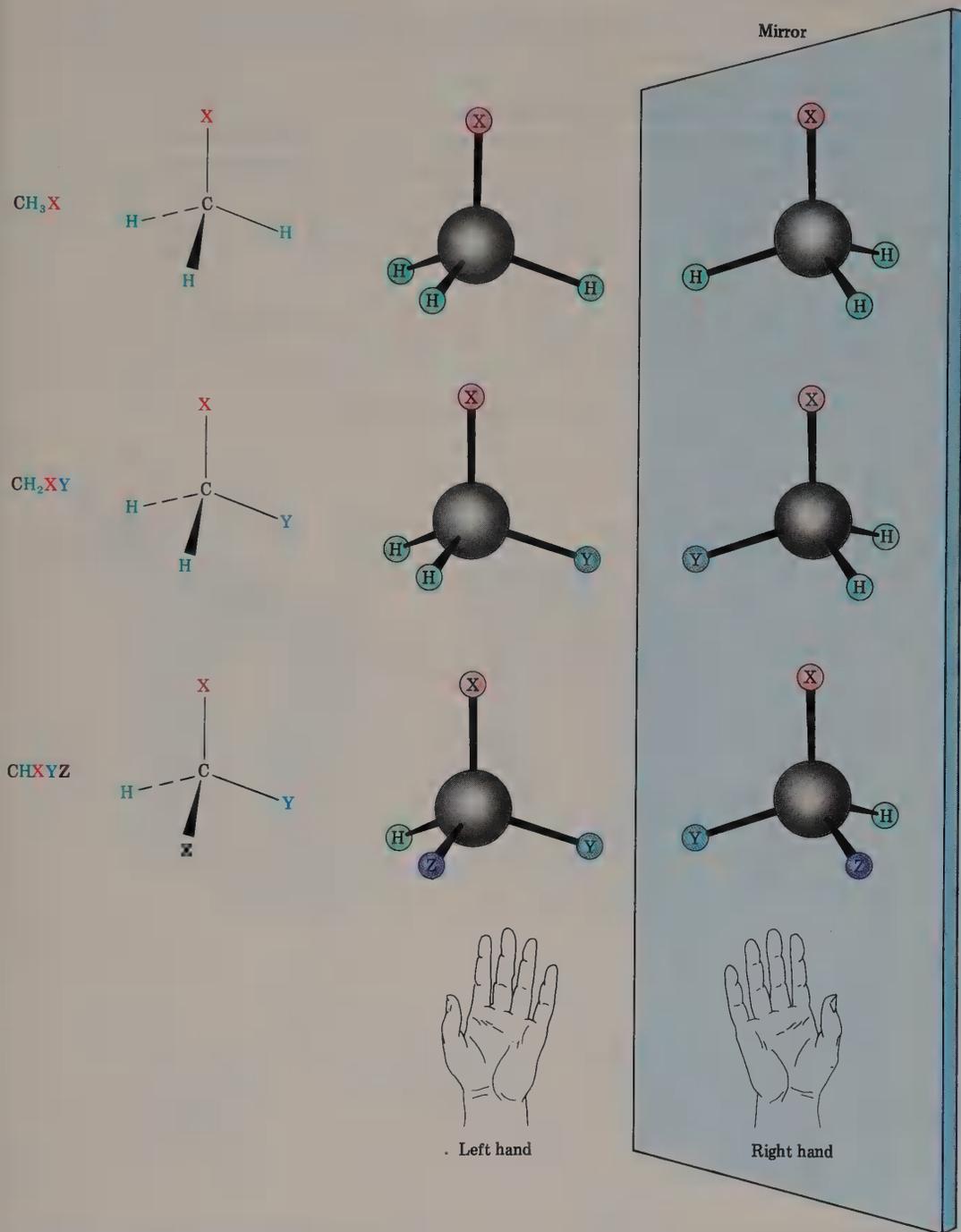
Handedness also plays a large role in organic chemistry as a direct consequence of the tetrahedral stereochemistry of  $sp^3$ -hybridized carbon. Let's see how handedness in organic molecules arises.

### 9.1 Enantiomers and the Tetrahedral Carbon

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Look at the generalized molecules of the type  $\text{CH}_3\text{X}$ ,  $\text{CH}_2\text{XY}$ , and  $\text{CHXYZ}$  shown in Figure 9.1. On the left are three molecules, and on the right are their images reflected in a mirror. The  $\text{CH}_3\text{X}$  and  $\text{CH}_2\text{XY}$  molecules are identical with their mirror images and thus are not handed. If you make molecular models of each molecule and of its mirror image, you can superimpose one on top of the other.

Unlike the  $\text{CH}_3\text{X}$  and  $\text{CH}_2\text{XY}$  molecules, though, the  $\text{CHXYZ}$  molecule is *not* identical with its mirror image. You can't superimpose a model of the molecule on top of a model of its mirror image for the same reason that you can't superimpose a left hand on a right hand. You might get *two* of the substituents superimposed, X and Y for example, but H and Z would be



**Figure 9.1** Tetrahedral carbon atoms and their mirror images. Molecules of the type  $\text{CH}_3\text{X}$  and  $\text{CH}_2\text{XY}$  are identical with their mirror images, but a molecule of the type  $\text{CHXYZ}$  is not. A  $\text{CHXYZ}$  molecule is related to its mirror image in the same way that a right hand is related to a left hand.

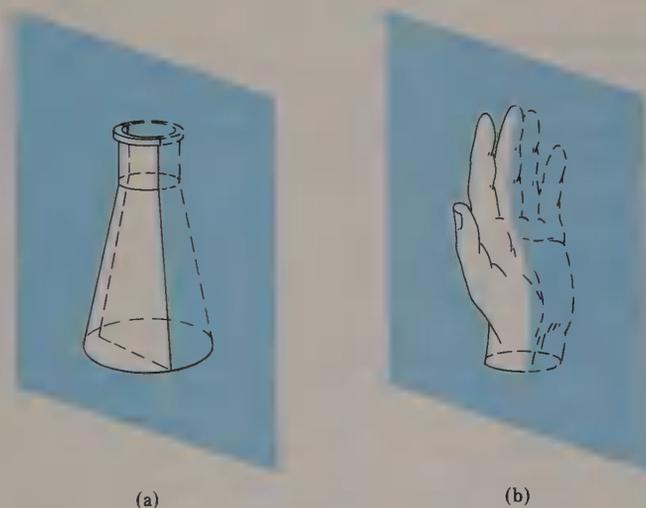


## 9.2 The Reason for Handedness in Molecules: Chirality

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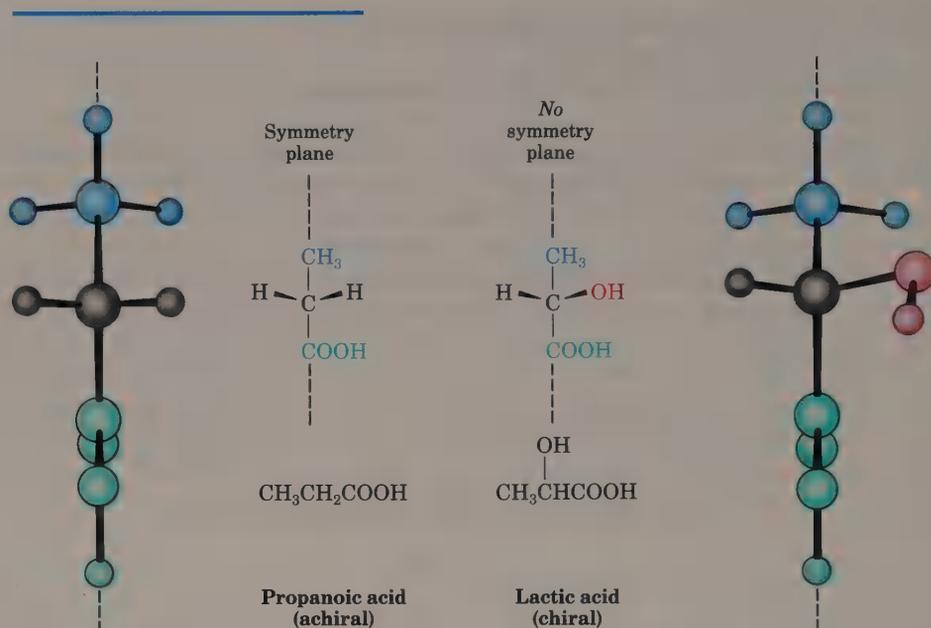
Molecules that are not identical with their mirror images and thus exist in two enantiomeric forms are said to be **chiral** (ky-ral, from the Greek *cheir*, “hand”). You can’t take a chiral molecule and its mirror image (enantiomer) and place one on top of the other so that all atoms coincide.

How can you predict whether a certain molecule is or is not chiral? *A molecule can’t be chiral if it contains a plane of symmetry.* A **plane of symmetry** is an imaginary plane that cuts through an object (or molecule) in such a way that one half of the object is an exact mirror image of the other half. For example, a flask has a plane of symmetry. If you were to cut the flask in half, one half would be an exact mirror image of the second half. A hand, however, has no plane of symmetry. One “half” of a hand is not a mirror image of the other half (Figure 9.3).



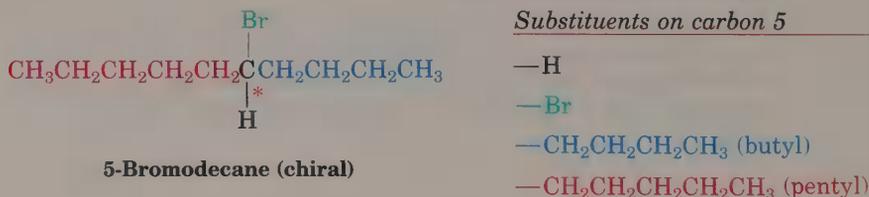
**Figure 9.3** The meaning of *symmetry plane*. An object like the flask (a) has a symmetry plane cutting through it, making right and left halves mirror images. An object like a hand (b) has no symmetry plane; the right “half” of a hand is not a mirror image of the left “half.”

A molecule that has a plane of symmetry in any of its possible conformations *must* be superimposable on its mirror image and hence must be nonchiral or **achiral** (a-ky-ral). Thus, propanoic acid has a plane of symmetry when it is lined up as shown in Figure 9.4 and is therefore achiral. Lactic acid (2-hydroxypropanoic acid), however, has no plane of symmetry and is thus chiral.



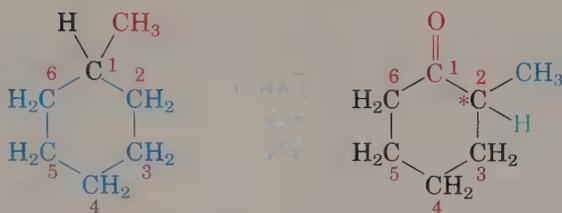
**Figure 9.4** The achiral propanoic acid molecule versus the chiral lactic acid molecule. Propanoic acid has a plane of symmetry that makes one side of the molecule a mirror image of the other side. Lactic acid, however, has no such symmetry plane.

The most common (although not the only) cause of chirality in organic molecules is the presence of a carbon atom bonded to four different groups, for example, the central carbon atom in lactic acid. Such carbons are referred to as *chiral centers*, or **stereogenic centers**. Detecting stereogenic centers in complex molecules takes practice because it's not always immediately apparent that four different groups are bonded to a given carbon. The differences don't necessarily appear right next to the stereogenic center. For example, 5-bromodecane is a chiral molecule because four different groups are bonded to C5, the stereogenic center (marked by an asterisk):



A butyl substituent is *similar* to a pentyl substituent, but it isn't identical. The difference isn't apparent until four carbon atoms away from the stereogenic center, but there's still a difference.

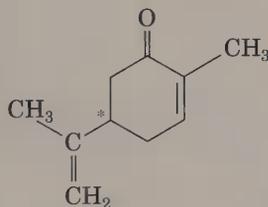
As other examples, look at methylcyclohexane and 2-methylcyclohexanone. Are either of these molecules chiral?



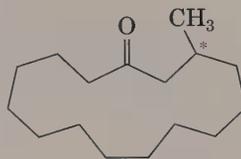
Methylcyclohexane (achiral)      2-Methylcyclohexanone (chiral)

Methylcyclohexane is achiral because no carbon atom in the molecule is bonded to four different groups. You can immediately eliminate all  $-\text{CH}_2-$  carbons and the  $-\text{CH}_3$  carbon from consideration, but what about C1 on the ring? The C1 carbon atom is bonded to a  $-\text{CH}_3$  group, to an  $-\text{H}$  atom, and to C2 and C6 of the ring. Carbons 2 and 6 are equivalent, however, as are carbons 3 and 5. Thus, the C6-C5-C4 “substituent” is equivalent to the C2-C3-C4 “substituent,” and methylcyclohexane is therefore achiral. An alternative way of arriving at the same conclusion is to realize that methylcyclohexane has a symmetry plane passing through the methyl group and through carbons 1 and 4 of the ring. Make a molecular model to see this symmetry plane more clearly.

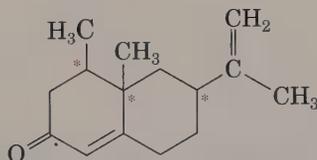
The situation is different for 2-methylcyclohexanone. 2-Methylcyclohexanone has no symmetry plane and is chiral because C2 is bonded to four different groups: a  $-\text{CH}_3$  group, an  $-\text{H}$  atom, a  $-\text{COCH}-$  ring bond (C1), and a  $-\text{CH}_2\text{CH}_2-$  ring bond (C3). Several more examples of chiral molecules are shown below. Check for yourself that the labeled centers are indeed chiral. [Remember:  $-\text{CH}_2-$ ,  $-\text{CH}_3$ ,  $\text{C}=\text{C}$ , and  $\text{C}\equiv\text{C}$  centers *can't* be chiral, because they have at least two identical bonds.]



Carvone (spear mint oil)



Muscone (musk oil)



Nootkatone (grapefruit oil)

PROBLEM.....

9.1 Which of these objects are chiral (handed)?

- (a) A screwdriver  
(d) A shoe

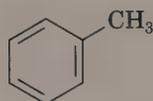
- (b) A screw  
(e) A hammer

- (c) A bean stalk

## PROBLEM.....

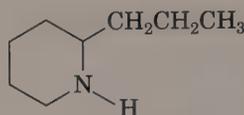
9.2 Which of these compounds are chiral? Build molecular models for help in seeing spatial relationships.

(a)

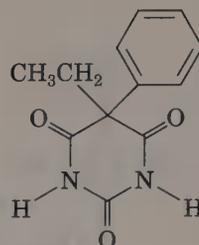


Toluene

(b)

Coniine  
(from poison hemlock)

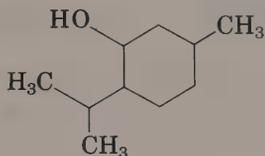
(c)

Phenobarbital  
(tranquilizer)

## PROBLEM.....

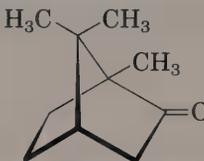
9.3 Place asterisks at all the stereogenic centers in these molecules:

(a)



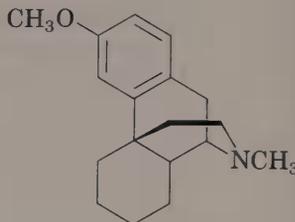
Menthol

(b)



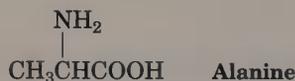
Camphor

(c)

Dextromethorphan  
(a cough suppressant)

## PROBLEM.....

9.4 Alanine, an amino acid found in proteins, is a chiral molecule. Use the standard convention of wedged, solid, and dashed lines to draw the two enantiomers of alanine.

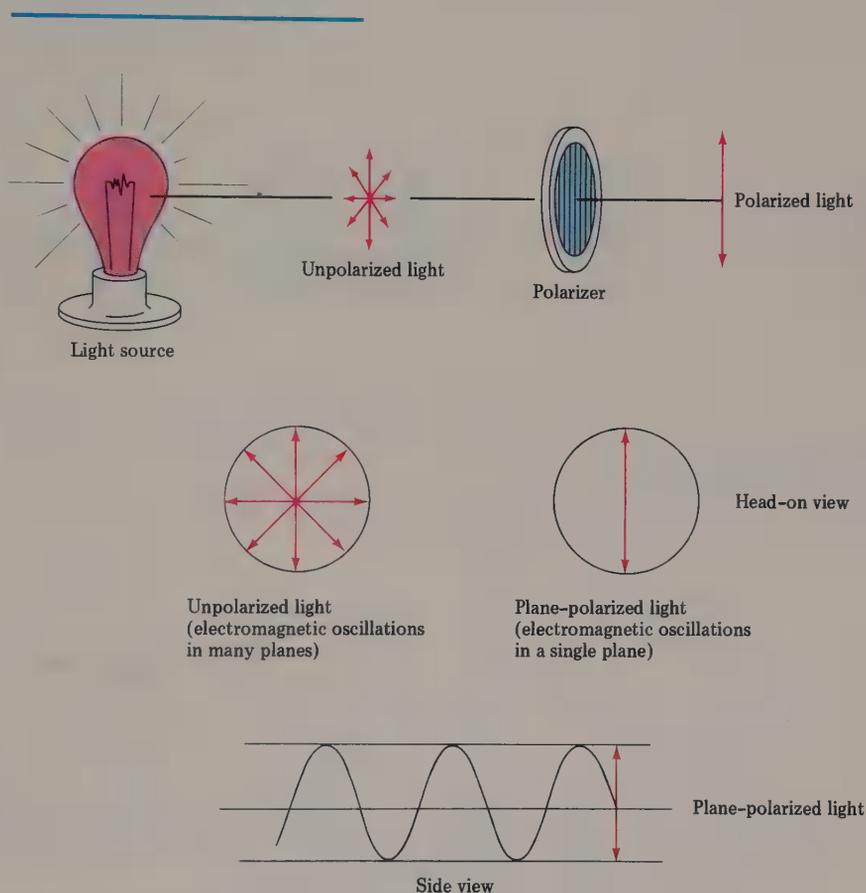


## 9.3 Optical Activity

The study of stereochemistry has its origins in the work of the French scientist Jean Baptiste Biot<sup>1</sup> in the early nineteenth century. Biot, a physicist, was investigating the nature of **plane-polarized light**. A beam of

<sup>1</sup>Jean Baptiste Biot (1774–1862); b. Paris; physicist, College de France.

ordinary light consists of electromagnetic waves that oscillate at right angles to the direction of light travel. Since ordinary light is unpolarized, this oscillation takes place in an infinite number of planes. When a beam of ordinary light is passed through a device called a *polarizer*, however, only the light waves oscillating in a *single* plane pass through; light waves in all other planes are blocked out. The light that passes through the polarizer has its electromagnetic waves vibrating in a well-defined plane, hence the name *plane-polarized light*. The polarization process is represented in Figure 9.5.

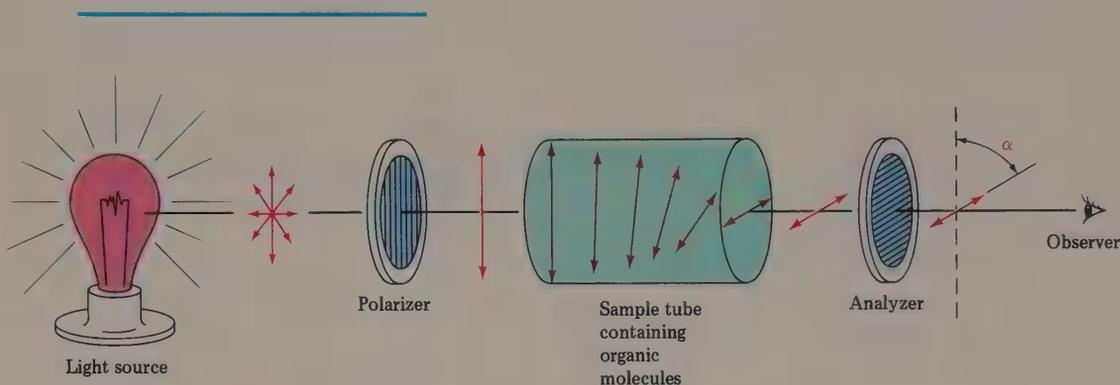


**Figure 9.5** Plane-polarized light. Oscillation of the electromagnetic field occurs in a single plane.

Biot made the remarkable observation that, when a beam of plane-polarized light is passed through solutions of certain organic molecules such as sugar or camphor, the plane of polarization is *rotated*. Not all organic

molecules exhibit this property, but those that do rotate plane-polarized light are said to be **optically active**.

The amount of rotation can be measured with an instrument known as a *polarimeter*, represented schematically in Figure 9.6. A solution of optically active organic molecules is first placed in a sample tube, plane-polarized light is passed through the tube, and rotation of the plane occurs. The light then goes through a second polarizer known as the *analyzer*. By rotating the analyzer until the light passes through it, we can find the new plane of polarization and can tell to what extent rotation has occurred. The amount of rotation is denoted  $\alpha$  (Greek alpha) and is expressed in degrees.



**Figure 9.6** Schematic representation of a polarimeter. Plane-polarized light passes through a solution of optically active molecules, which rotate the plane of polarization.

In addition to determining the extent of rotation, we can also find out its direction. From the vantage point of an observer looking directly end-on at the analyzer, some optically active molecules rotate polarized light to the left (counterclockwise) and are said to be **levorotatory**, whereas others rotate polarized light to the right (clockwise) and are said to be **dextrorotatory**. By convention, rotation to the left is given a minus sign ( $-$ ), and rotation to the right is given a plus sign ( $+$ ). For example, ( $-$ )-morphine is levorotatory, and ( $+$ )-sucrose is dextrorotatory.

## 9.4 Specific Rotation

Since the rotation of plane-polarized light is an intrinsic property of optically active organic molecules, it follows that the amount of rotation depends on the number of molecules that the light beam encounters. The more molecules

the light encounters, the greater the observed rotation. Thus, the amount of rotation observed depends on both sample concentration and sample path length. If we double the concentration of sample in a tube, the observed rotation doubles. Similarly, if we keep the sample concentration constant but double the length of the sample tube, the observed rotation is doubled. It also turns out that the amount of rotation depends on the wavelength of the light used.

To express the data in a meaningful way so that comparisons can be made, we have to choose standard conditions. By convention, the **specific rotation**,  $[\alpha]_D$ , of a compound is defined as the observed rotation when light of 589 nanometer ( $1 \text{ nm} = 10^{-9} \text{ m}$ ) wavelength is used with a sample path length  $l$  of 1 decimeter ( $1 \text{ dm} = 10 \text{ cm}$ ) and a sample concentration  $C$  of 1 g/mL. (Light of 589 nm, the so-called sodium D line, is the yellow light emitted from common sodium lamps.)

$$[\alpha]_D = \frac{\text{Observed rotation (degrees)}}{\text{Path length, } l \text{ (dm)} \times \text{Concentration, } C \text{ (g/mL)}} = \frac{\alpha}{l \times C}$$

When optical rotation data are expressed in this standard way, the specific rotation,  $[\alpha]_D$ , is a physical constant characteristic of each optically active compound. Some examples are listed in Table 9.1.

**TABLE 9.1** Specific Rotations of Some Organic Molecules

<i>Compound</i>	$[\alpha]_D$ (degrees)	<i>Compound</i>	$[\alpha]_D$ (degrees)
Camphor	+44.26	Penicillin V	+223
Morphine	-132	Monosodium glutamate	+25.5
Sucrose	+66.47	Benzene	0
Cholesterol	-31.5	Acetic acid	0

**PROBLEM** .....

- 9.5 Suppose you have a sucrose solution that appears to rotate plane-polarized light  $90^\circ$  to the right (dextrorotatory). How do you know for sure that the solution isn't rotating the plane of polarization to the *left* by  $270^\circ$  (levorotatory)? After all, the analyzer would be in exactly the same position in either case. [*Hint*: What effect would diluting the sample concentration have in each case?]

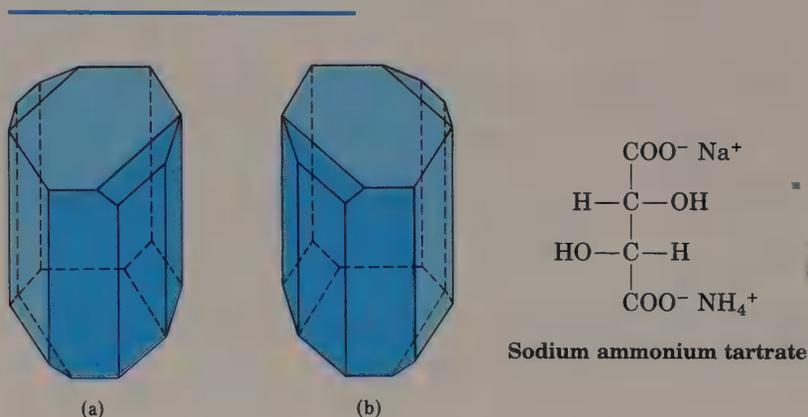
**PROBLEM** .....

- 9.6 A 1.5 g sample of coniine, the toxic extract of poison hemlock, was dissolved in 10 mL ethanol and placed in a sample cell with a 5.0 cm path length. The observed rotation at the sodium D line was  $+1.2^\circ$ . Calculate the specific rotation,  $[\alpha]_D$ , for coniine.
- .....

## 9.5 Pasteur's Discovery of Enantiomers

Little was done after Biot's discovery of optical activity in 1815 until Louis Pasteur<sup>2</sup> entered the picture in 1849. Pasteur had received his formal training in chemistry but had become interested in the subject of crystallography. He began work on crystalline salts of tartaric acid derived from wine and was repeating some measurements published a few years earlier when he made a surprising observation. When he recrystallized a concentrated solution of sodium ammonium tartrate below 28°C, two distinct kinds of crystals precipitated. Furthermore, the two kinds of crystals were nonsuperimposable mirror images of each other. That is, the crystals were not symmetrical but were related to each other in the same way that a right hand is related to a left hand.

Working carefully with a pair of tweezers, Pasteur was able to separate the crystals into two piles, one of "right-handed" crystals and one of "left-handed" crystals like those shown in Figure 9.7. Although a solution of the original salt (a 50:50 mixture of right and left) was optically inactive, *solutions of the crystals in each of the individual piles were optically active*, and their specific rotations were equal in amount but opposite in sign.



**Figure 9.7** Drawings of sodium ammonium tartrate crystals taken from Pasteur's original sketches. One of the crystals is "right-handed" and one is "left-handed."

Pasteur was far ahead of his time. Although the structural theory of Kekulé had not yet been proposed, Pasteur explained his results by speaking of the *molecules themselves*, saying: "There is no doubt that [in the *dextro* tartaric acid] there exists an asymmetric arrangement having a nonsuperimposable image. It is no less certain that the atoms of the *levo* acid possess

<sup>2</sup>Louis Pasteur (1822–1895); b. Dôle, Jura, France; studied at Arbois, Besançon; professor, Dijon, Strasbourg (1849–1854), Lille (1854–1857), École Normale Supérieure (1857–1863).

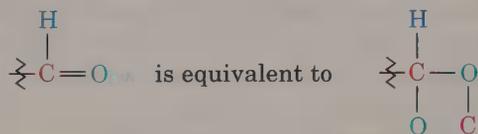
precisely the inverse asymmetric arrangement." Pasteur's vision was extraordinary, for it was not until 25 years later that the theories of van't Hoff and Le Bel confirmed his ideas regarding the asymmetric carbon atom.

Today, of course, we would describe Pasteur's work by saying that he had discovered the phenomenon of enantiomerism. The enantiomeric tartaric acid salts that Pasteur separated are identical in all respects except for their interaction with plane-polarized light. They have the same melting point, the same boiling point, the same solubilities, and the same spectroscopic properties.

## 9.6 Sequence Rules for Specification of Configuration

Although drawings provide a pictorial representation of stereochemistry, they are difficult to translate into words. Thus, a verbal method for indicating the three-dimensional arrangement of atoms (the **configuration**) at a stereogenic center is also necessary. The method used employs the same sequence rules that we used for the specification of alkene geometry (*Z* versus *E*) in Section 6.6. Let's briefly review the Cahn–Ingold–Prelog sequence rules to see how they can be used for specifying the configuration of a stereogenic center. Refer to Section 6.6 for an explanation of each rule.

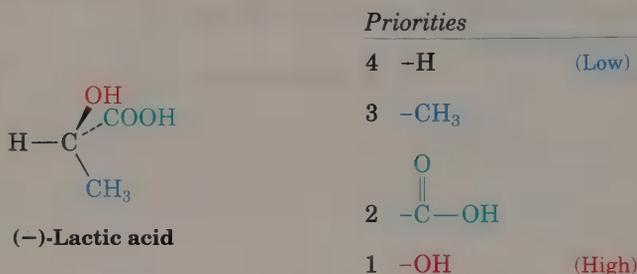
1. Look at the four atoms directly attached to the stereogenic center and assign priorities in order of decreasing atomic number. The group with highest atomic number is ranked first; the group with lowest atomic number is ranked fourth.
2. If a decision about priority can't be reached by applying rule 1, compare atomic numbers of the second atoms in each substituent, continuing on as necessary through the third or fourth atoms outward until the first point of difference is reached.
3. Multiple-bonded atoms are considered as an equivalent number of single-bonded atoms. For example:



Having assigned priorities to the four groups attached to a chiral carbon, we describe the stereochemical configuration around the chiral carbon by mentally orienting the molecule so that the group of lowest priority (fourth) is pointing directly back, away from us. We then look at the three remaining substituents, which now appear to radiate toward us like the spokes on a steering wheel (Figure 9.8). If a curved arrow drawn from the highest to second-highest to third-highest priority substituent (1  $\rightarrow$  2  $\rightarrow$  3) is *clockwise*, we say that the stereogenic center has the *R* configuration (Latin *rectus*,

“right”).<sup>3</sup> If a curved arrow from the highest to second-highest to third-highest priority substituent (1 → 2 → 3) is *counterclockwise*, the stereogenic center has the *S* configuration (Latin *sinister*, “left”). To remember these assignments, think of a car’s steering wheel when making a right (clockwise) or left (counterclockwise) turn.

Look at (–)-lactic acid to see an example of how configuration is assigned. The first step is to assign priorities to the four substituents. Sequence rule 1 says that –OH is first priority and –H is fourth priority, but doesn’t allow us to distinguish between –CH<sub>3</sub> and –COOH, since both groups have carbon as their first atom. Sequence rule 2 says that –COOH is higher priority than –CH<sub>3</sub>, since oxygen outranks hydrogen (the second atom in each group).

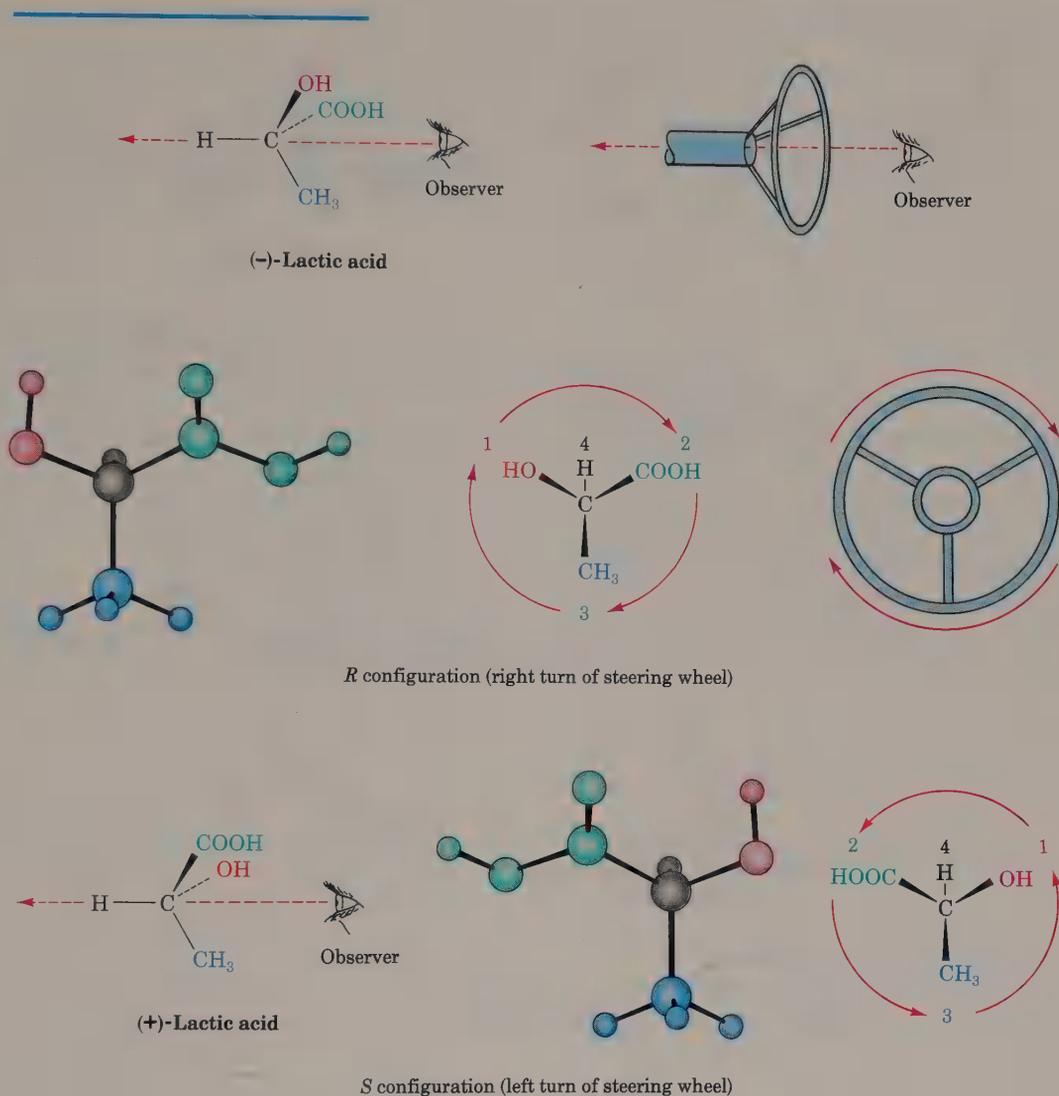


The next step is to turn the molecule so that the fourth-priority group (–H) is oriented toward the rear, away from the observer. Since a curved arrow from 1 (–OH) to 2 (–COOH) to 3 (–CH<sub>3</sub>) is clockwise (right turn of the steering wheel), we assign the *R* configuration to (–)-lactic acid. Applying the same procedure to (+)-lactic acid should (and does) lead to the opposite assignment, as shown in Figure 9.8.

Further examples of how to assign configuration are provided by naturally occurring (+)-alanine and (–)-glyceraldehyde, which have the *S* configurations shown in Figure 9.9. *Note that the sign of optical rotation, (+) or (–), is not related to the R,S designation.* (*S*)-Alanine happens to be dextrorotatory (+), and (*S*)-glyceraldehyde happens to be levorotatory (–). There is no simple correlation between *R,S* configuration and direction or magnitude of optical rotation.

One further point needs mentioning: the matter of **absolute configuration**. How do we know that our assignments of *R,S* configuration are correct in an *absolute*, rather than a relative, sense? Since we can’t see the molecules themselves, how do we know for certain that it is the dextrorotatory enantiomer of lactic acid that has the *R* configuration? This difficult question was not solved until 1951 when J. M. Bijvoet of the University of Utrecht reported an X-ray spectroscopic method for determining the absolute spatial arrangement of atoms in a molecule. Based on his results, we can say with certainty that the *R,S* conventions are correct.

<sup>3</sup>*Rectus* actually means “right” in the sense of “correct,” whereas *dexter* is the Latin word for “right” in the directional sense. In some European societies, however, left-handedness has historically been looked on with disfavor, and it has been considered “right” to be right-handed.

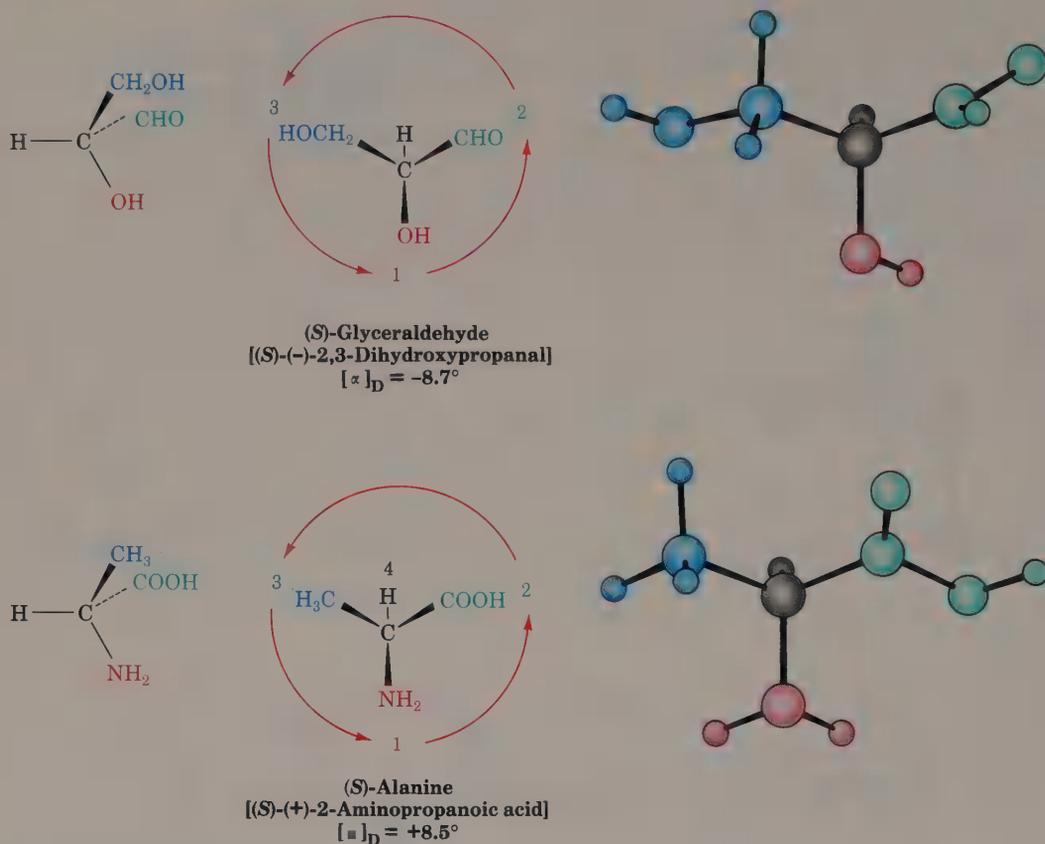


**Figure 9.8** Assignment of configuration to (*R*)-(-)-lactic acid and (*S*)-(+)-lactic acid.

**PRACTICE PROBLEM**.....

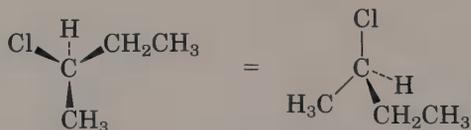
Draw a tetrahedral representation of *R*-2-chlorobutane.

**Solution** The four substituents bound to the chiral carbon of *R*-2-chlorobutane can be assigned the following priorities: (1)  $-\text{Cl}$ , (2)  $-\text{CH}_2\text{CH}_3$ , (3)  $-\text{CH}_3$ , (4)  $-\text{H}$ . To draw a tetrahedral representation of the molecule, first orient the low-priority  $-\text{H}$  group toward the rear and imagine that the other three groups are coming out of the page toward you. Place the remaining three substituents in order such that the direction of travel from  $1 \rightarrow 2 \rightarrow 3$  is clockwise (right turn), and then tilt the molecule slightly



**Figure 9.9** Assignment of configuration to (-)-glyceraldehyde and (+)-alanine. Both happen to have the *S* configuration.

to bring the rear hydrogen into view and end up with a standard tetrahedral representation:



Using molecular models is a great help in working problems of this sort.

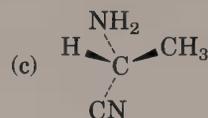
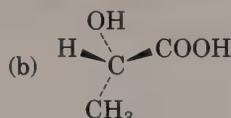
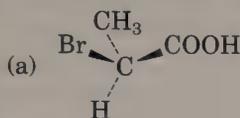
**PROBLEM** .....

**9.7** Assign priorities to these sets of substituents:

- H, -Br, -CH<sub>2</sub>CH<sub>3</sub>, -CH<sub>2</sub>CH<sub>2</sub>OH
- CO<sub>2</sub>H, -CO<sub>2</sub>CH<sub>3</sub>, -CH<sub>2</sub>OH, -OH
- CN, -CH<sub>2</sub>NH<sub>2</sub>, -CH<sub>2</sub>NHCH<sub>3</sub>, -NH<sub>2</sub>
- Br, -CH<sub>2</sub>Br, -Cl, -CH<sub>2</sub>Cl

## PROBLEM

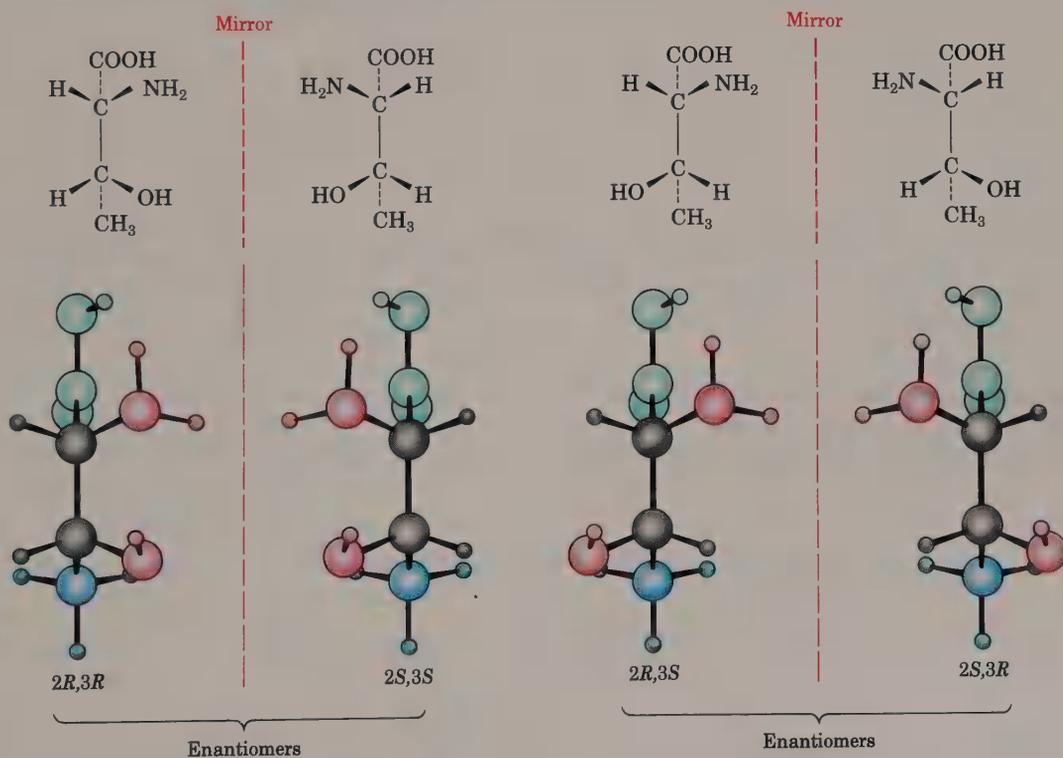
9.8 Assign *R,S* configurations to these molecules:



## 9.7 Diastereomers

Molecules like lactic acid, alanine, and glyceraldehyde are relatively simple to deal with, since each has only one stereogenic center and can exist in only two enantiomeric configurations. The situation becomes more complex, however, with molecules that have more than one stereogenic center.

Look at the essential amino acid threonine (2-amino-3-hydroxybutanoic acid), for example. Since threonine has two stereogenic centers (C2 and C3), there are four possible stereoisomers, as shown in Figure 9.10. Check for yourself that the *R,S* configurations are correct as indicated.



**Figure 9.10** The four stereoisomers of 2-amino-3-hydroxybutanoic acid (threonine).

The four threonine stereoisomers can be classified into two mirror-image pairs of enantiomers. The  $2R,3R$  stereoisomer is the mirror image of  $2S,3S$ , and the  $2R,3S$  stereoisomer is the mirror image of  $2S,3R$ . But what is the relationship between any two configurations that are not mirror images? What, for example, is the relationship between the  $2R,3R$  compound and the  $2R,3S$  compound? The two compounds are stereoisomers, yet they aren't superimposable and they aren't enantiomers. To describe such a relationship, we need a new term.

**Diastereomers** are stereoisomers that are not mirror images of each other. Since we used the right-hand/left-hand analogy to describe the relationship between two enantiomers, we might extend the analogy further by saying that diastereomers have a hand-foot relationship. Hands and feet look very *similar*, but they aren't identical and they aren't mirror images. The same is true of diastereomers; they're similar but not identical. Note carefully the difference between enantiomers and diastereomers: Enantiomers must have opposite (mirror-image) configurations at *all* stereogenic centers; diastereomers must have opposite configurations at *some* (one or more) stereogenic centers, but the same configuration at others. A full description of the four threonine stereoisomers is given in Table 9.2.

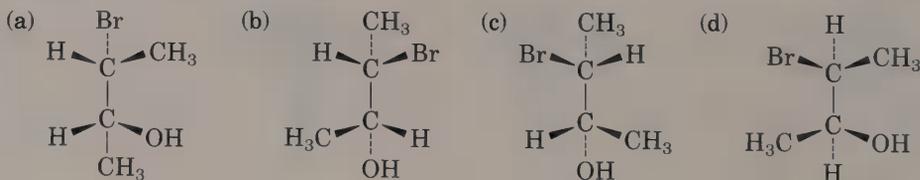
**TABLE 9.2** Relationships Between Four Stereoisomeric Threonines

<i>Stereoisomer</i>	<i>Enantiomeric with</i>	<i>Diastereomeric with</i>
$2R,3R$	$2S,3S$	$2R,3S$ and $2S,3R$
$2S,3S$	$2R,3R$	$2R,3S$ and $2S,3R$
$2R,3S$	$2S,3R$	$2R,3R$ and $2S,3S$
$2S,3R$	$2R,3S$	$2R,3R$ and $2S,3S$

Of the four possible stereoisomers of threonine, only the  $2S,3R$  isomer,  $[\alpha]_D = -29.3^\circ$ , occurs naturally in plants and animals. Most biologically important molecules are chiral, and usually only a single stereoisomer is found in nature.

**PROBLEM** .....

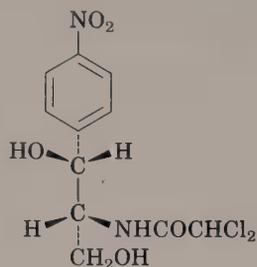
- 9.9** Assign  $R,S$  configurations to these molecules. Which are enantiomers, and which are diastereomers?



**PROBLEM** .....

- 9.10** Chloramphenicol, a powerful antibiotic isolated in 1949 from the *Streptomyces venezuelae* bacterium, is active against a broad spectrum of bacterial infections and is

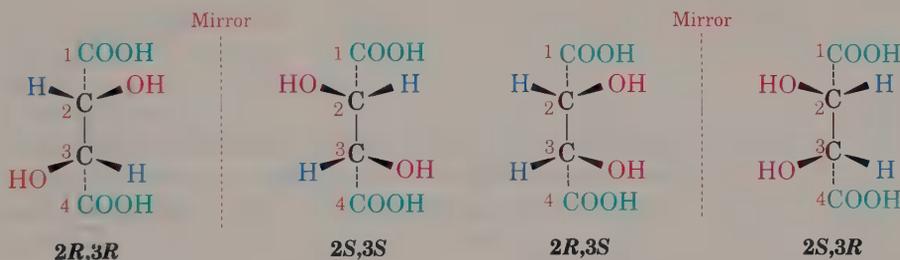
particularly valuable against typhoid fever. Assign *R,S* configurations to the stereogenic centers in chloramphenicol.



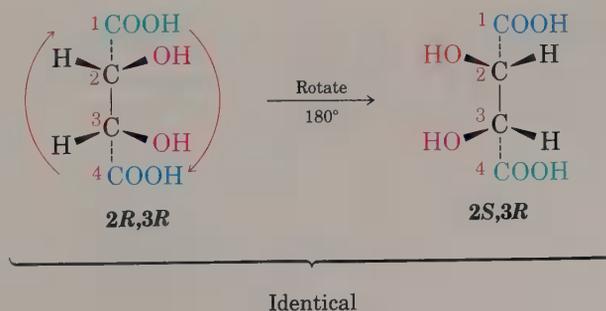
**Chloramphenicol**  
 $[\alpha]_D = +18.6^\circ$

## 9.8 Meso Compounds

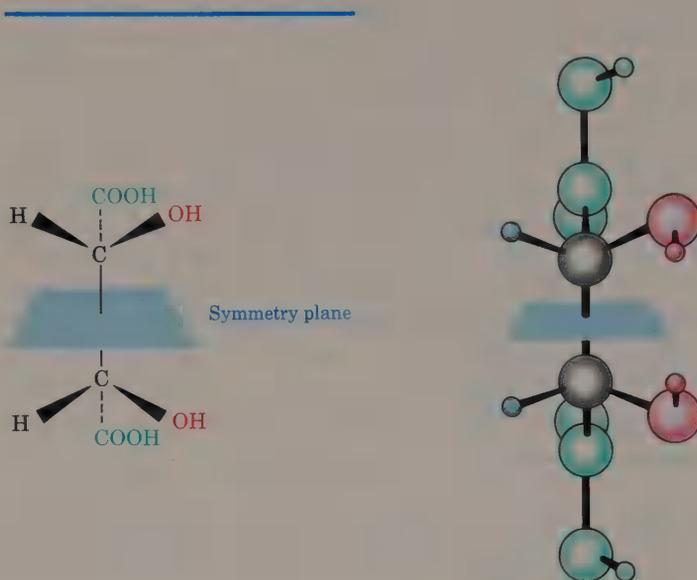
Let's look at one more example of a compound with two stereogenic centers: tartaric acid. We're already acquainted with tartaric acid for its role in Pasteur's discovery of optical activity, and we can now draw the four stereoisomers:



The mirror-image *2R,3R* and *2S,3S* structures are not superimposable and are therefore a pair of enantiomers. A close look, however, reveals that the *2R,3S* and *2S,3R* structures are *identical*, as we can see by rotating one structure  $180^\circ$ :



The identity of the  $2R,3S$  and  $2S,3R$  structures results from the fact that the molecule has a plane of symmetry. The symmetry plane cuts through the C2–C3 bond, making one half of the molecule a mirror image of the other half (Figure 9.11).



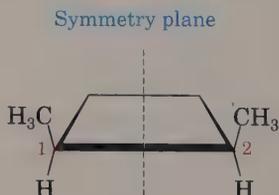
**Figure 9.11** A symmetry plane through the C2–C3 bond of *meso*-tartaric acid.

Because of the plane of symmetry, the tartaric acid stereoisomer shown in Figure 9.11 must be achiral, despite the fact that it has two stereogenic centers. Compounds that are achiral by virtue of a symmetry plane, yet contain stereogenic centers, are called **meso compounds** (*me-zo*). Thus, tartaric acid exists in three stereoisomeric configurations: two enantiomers and one meso form.

**PRACTICE PROBLEM**.....

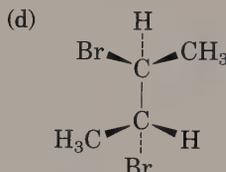
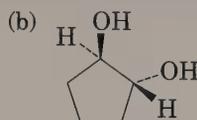
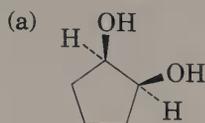
Does *cis*-1,2-dimethylcyclobutane have any stereogenic centers? Is it a chiral molecule?

**Solution** Looking at the structure of *cis*-1,2-dimethylcyclobutane, we see that both of the methyl-bearing ring carbons (C1 and C2) are chiral. Overall, though, the compound is achiral because there is a symmetry plane bisecting the ring between C1 and C2. Thus, the molecule is a meso compound.



## PROBLEM

9.11 Which, if any, of these structures represent meso compounds?



## PROBLEM

9.12 Which of these substances have a meso form?

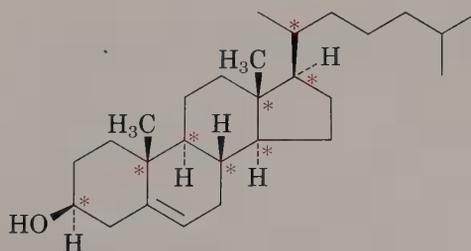
(a) 2,3-Dibromobutane

(b) 2,3-Dibromopentane

(c) 2,4-Dibromopentane

## 9.9 Molecules with More Than Two Stereogenic Centers

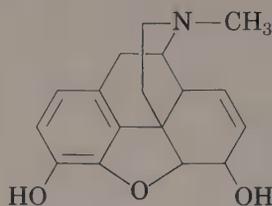
We've seen how a single stereogenic center in a molecule gives rise to two stereoisomers (one pair of enantiomers) and how two stereogenic centers in a molecule give rise to a maximum of four stereoisomers (two pairs of enantiomers). In general, a molecule with  $n$  stereogenic centers gives rise to a maximum of  $2^n$  stereoisomers ( $2^{n-1}$  pairs of enantiomers). For example, cholesterol contains eight stereogenic centers making possible  $2^8 = 256$  stereoisomers (128 enantiomeric pairs), although many would be too highly strained to exist. Only one, however, is produced in nature.



**Cholesterol**  
(eight chiral centers)

## PROBLEM .....

- 9.13 How many stereogenic centers does morphine have? How many stereoisomers of morphine are possible in principle?



Morphine

## 9.10 Racemic Mixtures

To conclude our discussion of stereoisomerism, let's return for a final look at Pasteur's pioneering work. Pasteur took an optically inactive form of a tartaric acid salt and found that he could crystallize two optically active forms from it. These two optically active forms were the *2R,3R* and *2S,3S* configurations just discussed, but what was the optically inactive form he started with? It couldn't have been *meso*-tartaric acid, since *meso*-tartaric acid is a different chemical compound from the two chiral enantiomers and can't interconvert with them without breaking and re-forming chemical bonds.

The answer is that Pasteur started with a 50:50 mixture of the two chiral enantiomers of tartaric acid. Such a mixture is called a **racemic** (ray-see-mic) **mixture**, or **racemate**, and is denoted either by the symbol ( $\pm$ ) or by the prefix *d,l* to indicate a mixture of dextrorotatory and levorotatory forms. Racemic mixtures *must* show zero optical rotation because equal amounts of (+) and (−) forms are present. The (+) rotation from one enantiomer exactly cancels the (−) rotation from the other enantiomer. Through good fortune, Pasteur was able to **resolve** ( $\pm$ )-tartaric acid into its (−) and (+) enantiomers. Unfortunately, the fractional crystallization method he used doesn't work for most racemic mixtures, and other techniques are required. The method of resolution most often used is discussed in Section 25.5.

## 9.11 Physical Properties of Stereoisomers

If seemingly simple compounds such as tartaric acid can exist in several different stereoisomeric configurations, the question arises whether the different stereoisomers of a compound have different physical properties. The answer is yes, they do.

Some physical properties of the three different stereoisomers of tartaric acid and of the racemic mixture are listed in Table 9.3. As indicated, the (+)- and (−)-tartaric acids have identical melting points, solubilities, and

TABLE 9.3 Some Properties of the Stereoisomers of Tartaric Acid

Stereoisomer	Melting point (°C)	$[\alpha]_D$ (degrees)	Density (g/cm <sup>3</sup> )	Solubility at 20°C (g/100 mL H <sub>2</sub> O)
(+)	168–170	+12	1.7598	139.0
(-)	168–170	-12	1.7598	139.0
Meso	146–148	0	1.6660	125.0
(±)	206	0	1.7880	20.6

densities. They differ only in the sign of their rotation of plane-polarized light. The meso isomer, by contrast, is diastereomeric with the (+) and (-) forms. As such, it has no mirror-image relationship to (+)- and (-)-tartaric acids; it is a different compound altogether and has different physical properties.

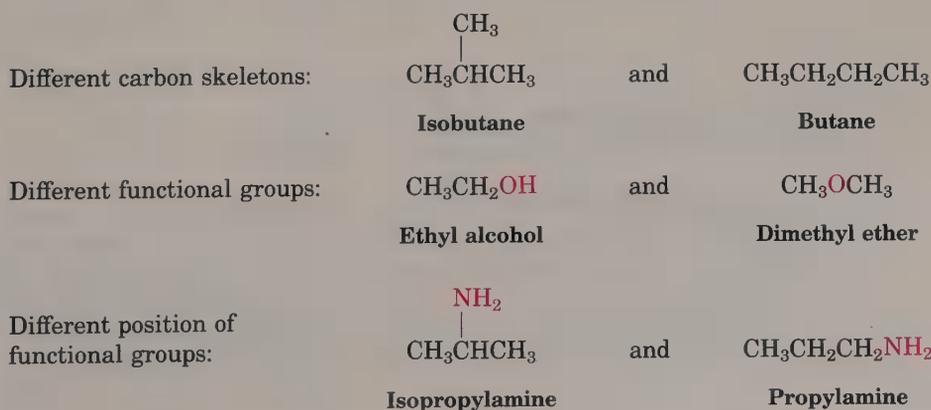
The racemic mixture is different still. Though a mixture of enantiomers, racemates act as though they were pure compounds, different from either enantiomer. As the table shows, the physical properties of racemic tartaric acid differ from those of the two pure enantiomers and from those of the meso form.

## 9.12 A Brief Review of Isomerism

We've seen several different kinds of isomers in the past few chapters, and it's a good idea at this point to see how they relate to one another. As noted earlier, isomers are compounds that have the same chemical formula but have different structures. There are two fundamental types of isomerism, both of which we've now encountered: constitutional isomerism and stereoisomerism.

*Constitutional isomers* (Section 3.2) are compounds whose atoms are connected differently. Among the different kinds of constitutional isomers we've seen are skeletal, functional, and positional isomers.

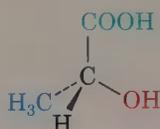
**Constitutional isomers**—different connections among atoms:



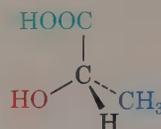
*Stereoisomers* (Section 3.9) are compounds that have their atoms connected in the same order but have a different geometry. Among the different kinds of stereoisomers we've seen are enantiomers, diastereomers, and cis–trans isomers (both in alkenes and in cycloalkanes). To be perfectly accurate, however, cis–trans isomers are really just a special category of diastereomers, since they meet the definition of being non-mirror-image stereoisomers:

**Stereoisomers**—same connections among atoms, but different geometry:

**Enantiomers**  
(nonsuperimposable  
mirror-image  
stereoisomers)



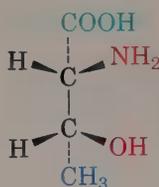
*(R)*-Lactic acid



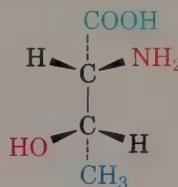
*(S)*-Lactic acid

**Diastereomers**  
(nonsuperimposable,  
non-mirror-image  
stereoisomers)

Configurational  
diastereomers

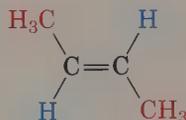


*2R,3R*-2-Amino-3-  
hydroxybutanoic acid



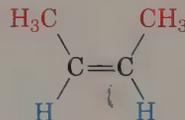
*2R,3S*-2-Amino-3-  
hydroxybutanoic acid

Cis–trans diastereomers  
(substituents on same  
side or opposite side of  
double bond or ring)



*trans*-2-Butene

and



*cis*-2-Butene



*trans*-1,3-  
Dimethylcyclopentane

and



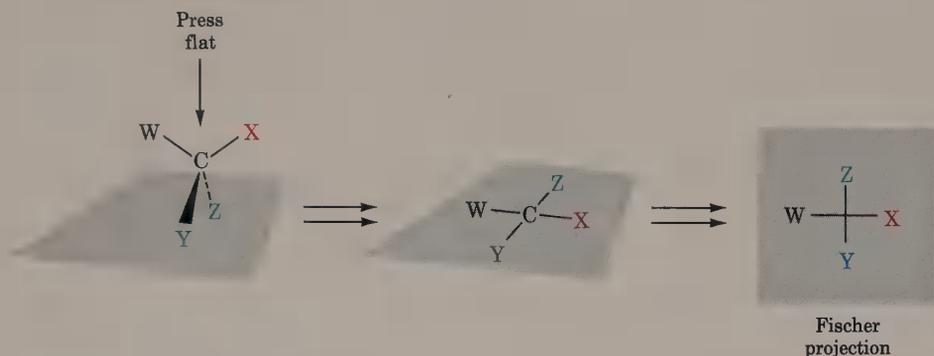
*cis*-1,3-  
Dimethylcyclopentane

## 9.13 Fischer Projections

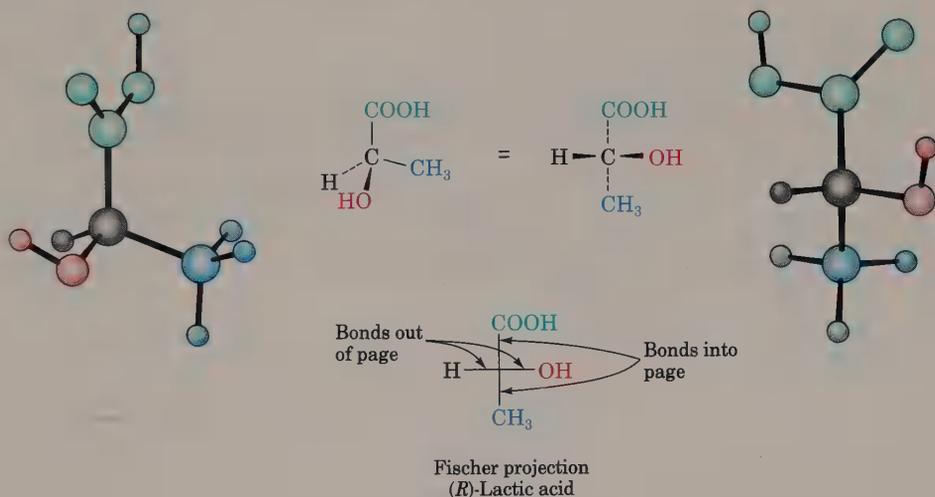
When learning to visualize chiral molecules, it's best to begin by building molecular models. As more experience is gained, it becomes easier to draw pictures and work with mental images. To do this successfully, though, a standard method of representation is needed for depicting the three-dimensional arrangement of atoms on a flat page. In 1891, Emil Fischer suggested a method based on the projection of a tetrahedral carbon atom

onto a flat surface. These **Fischer projections** were soon adopted and are now a standard means of depicting stereochemistry at stereogenic centers.

A tetrahedral carbon atom is represented in Fischer projection by two crossed lines. By convention, the horizontal lines represent bonds coming out of the page, and the vertical lines represent bonds going into the page:



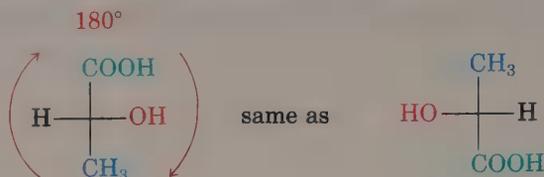
For example, (*R*)-lactic acid can be drawn as follows:



Since a given molecule can be depicted in many different ways, it's often necessary to compare two different projections to see if they represent the same or different enantiomers. To test for identity, Fischer projections can be moved around on the paper, but care must be taken not to change the meaning of the projection inadvertently. Only two kinds of motions are allowed:

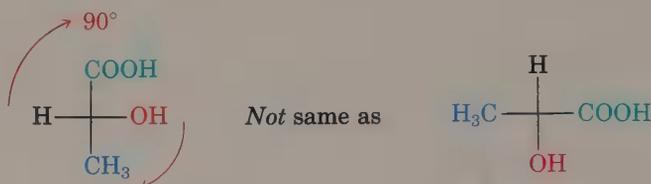
1. A Fischer projection can be rotated on the page by  $180^\circ$ , but *not* by  $90^\circ$  or  $270^\circ$ . The reason for this rule is simply that a  $180^\circ$  rotation obeys the Fischer convention by keeping the same two substituents on horizontal bonds always coming out of the plane and the same two substituents on vertical bonds always going into the plane. A  $90^\circ$  rotation, however, disobeys the Fischer convention. Substituents

that were on horizontal bonds in one form become vertical after a  $90^\circ$  rotation, thus inverting their stereochemistry and changing the meaning of the projection. For example:



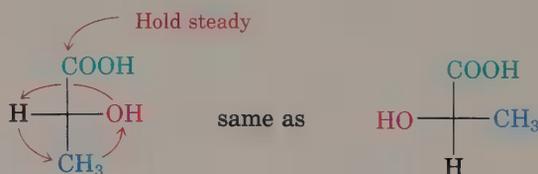
$-\text{COOH}$  and  $-\text{CH}_3$  go into plane of paper in both projections;  $-\text{H}$  and  $-\text{OH}$  come out of plane of paper in both projections.

but:



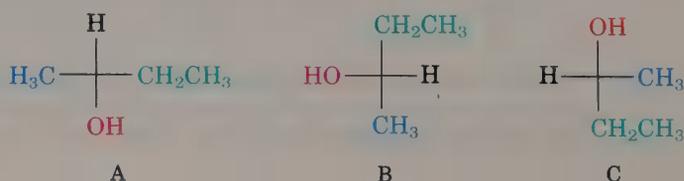
$-\text{COOH}$  and  $-\text{CH}_3$  groups go into plane of paper in one projection but come out of plane of paper in other projection.

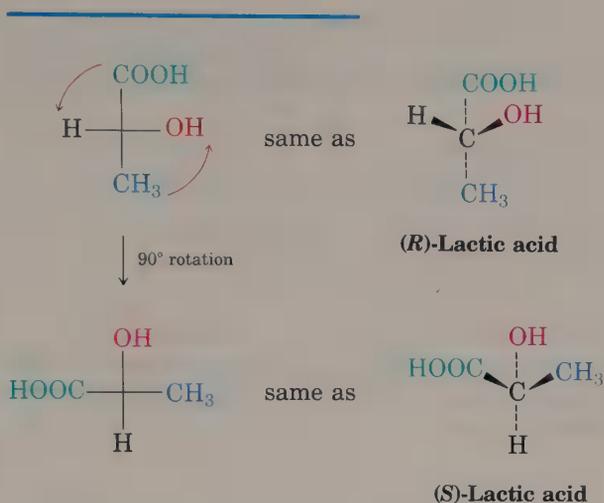
2. A Fischer projection can have one group held steady while the other three rotate in either a clockwise or a counterclockwise direction. For example:



These are the only kinds of motion allowed. Moving a Fischer projection in any other way inverts its meaning. Thus, if a Fischer projection of (*R*)-lactic acid is turned by  $90^\circ$ , a projection of (*S*)-lactic acid results (Figure 9.12).

Knowing the two rules provides an easy way to superimpose molecules mentally. For example, three different Fischer projections of 2-butanol follow. Do they all represent the same enantiomer, or is one different?

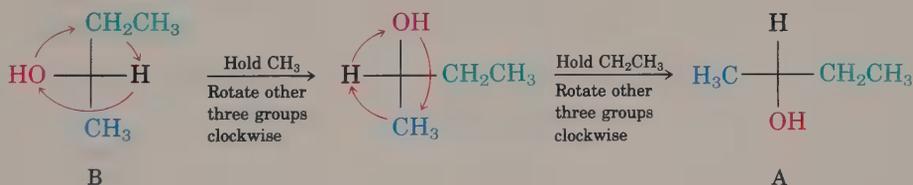




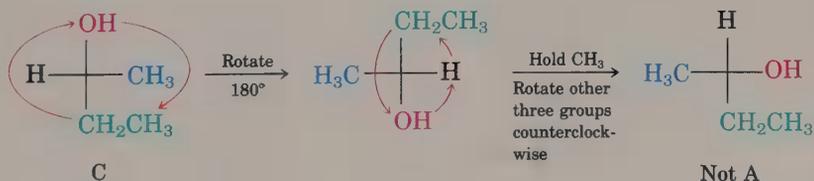
**Figure 9.12** Rotation of a Fischer projection by  $90^\circ$  inverts its meaning.

The simplest way to determine if two Fischer projections represent the same enantiomer is to carry out allowed motions until *two* groups are superimposed. If the other two groups are also superimposed, then the Fischer projections are the same; if the other two groups are not superimposed, the Fischer projections are different.

Let's keep projection A unchanged and move B so that the  $-\text{CH}_3$  and  $-\text{H}$  substituents match up with those in A:



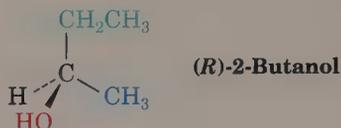
By performing two allowed movements on projection B, we find that it is identical to projection A. Now let's do the same thing to projection C:



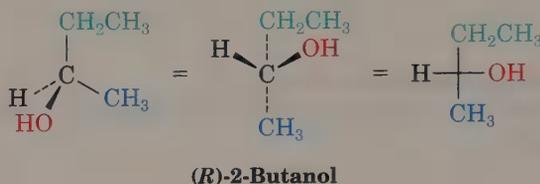
By performing two allowed movements on projection C, we can match up the  $-\text{H}$  and  $-\text{CH}_3$  substituents with those in A. After doing so, however, the  $-\text{OH}$  and  $-\text{CH}_2\text{CH}_3$  substituents *don't* match up. Thus, projection C must be enantiomeric with A and B.

## PRACTICE PROBLEM.....

Convert the following tetrahedral representation of (*R*)-2-butanol into a Fischer projection.

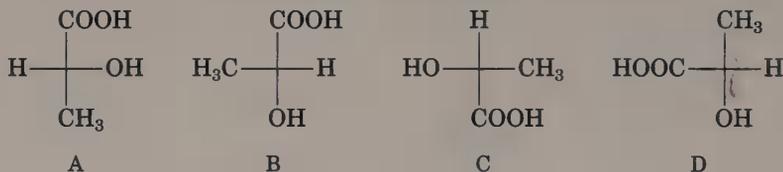


**Solution** Orient the molecule so that two horizontal bonds are facing you and two vertical bonds are receding from you. Then press the molecule flat into the paper, indicating the chiral carbon as the intersection of two crossed lines:



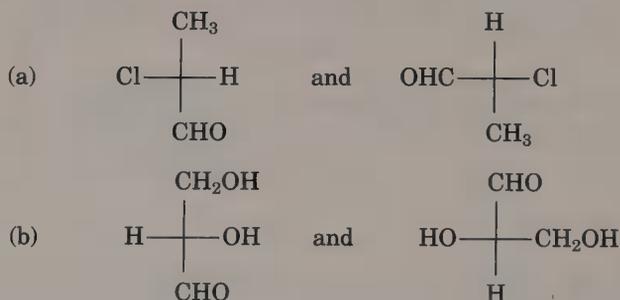
## PROBLEM.....

9.14 Which of these Fischer projections of lactic acid represent the same enantiomer?



## PROBLEM.....

9.15 Are these pairs of Fischer projections the same, or are they enantiomers?



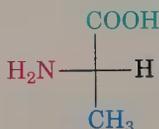
## 9.14 Assigning *R,S* Configurations to Fischer Projections

*R,S* stereochemical designations (Section 9.6) can be assigned to Fischer projections by following three steps:

1. Assign priorities to the four substituents in the usual way.
2. Perform one of the two allowed motions to place the group of lowest (fourth) priority at the top of the Fischer projection.
3. Determine the direction of rotation in going from priority 1 to 2 to 3, and assign *R* or *S* configuration as in the following practice problem.

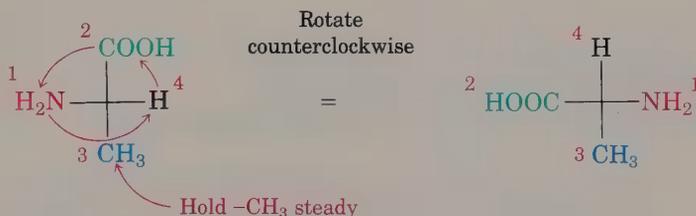
### PRACTICE PROBLEM.....

Assign *R* or *S* configuration to this Fischer projection of alanine, an amino acid:

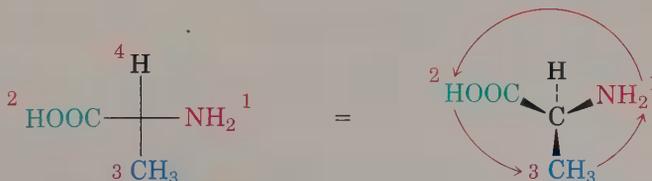


Alanine

**Solution** First, assign priorities to the four substituents on the chiral carbon. According to the sequence rules, the priorities are  $-\text{NH}_2$  (1),  $-\text{COOH}$  (2),  $-\text{CH}_3$  (3), and  $-\text{H}$  (4). Next, perform one of the allowed motions on the Fischer projection to bring the group of lowest priority ( $-\text{H}$ ) to the top. In the present instance, we might want to hold the  $-\text{CH}_3$  group steady while rotating the other three groups counterclockwise:

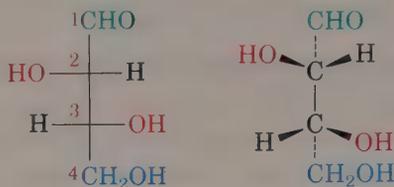


Going now from first to second to third highest priority requires a counterclockwise (left-hand) turn, corresponding to *S* stereochemistry.



*S* stereochemistry

Fischer projections can be used to specify more than one stereogenic center in a molecule simply by “stacking” the stereogenic centers, one on top of the other. For example, threose, a simple four-carbon sugar, has the  $2S,3R$  configuration:

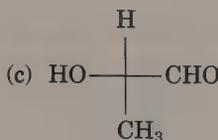
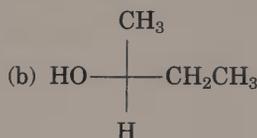
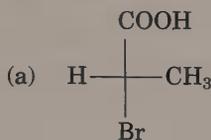


Threose [( $2S,3R$ )-2,3,4-Trihydroxybutanal]

Molecular models are particularly helpful in visualizing these structures.

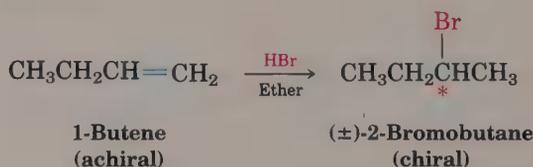
PROBLEM .....

9.16 Assign either  $R$  or  $S$  configuration to the stereogenic centers in these molecules:

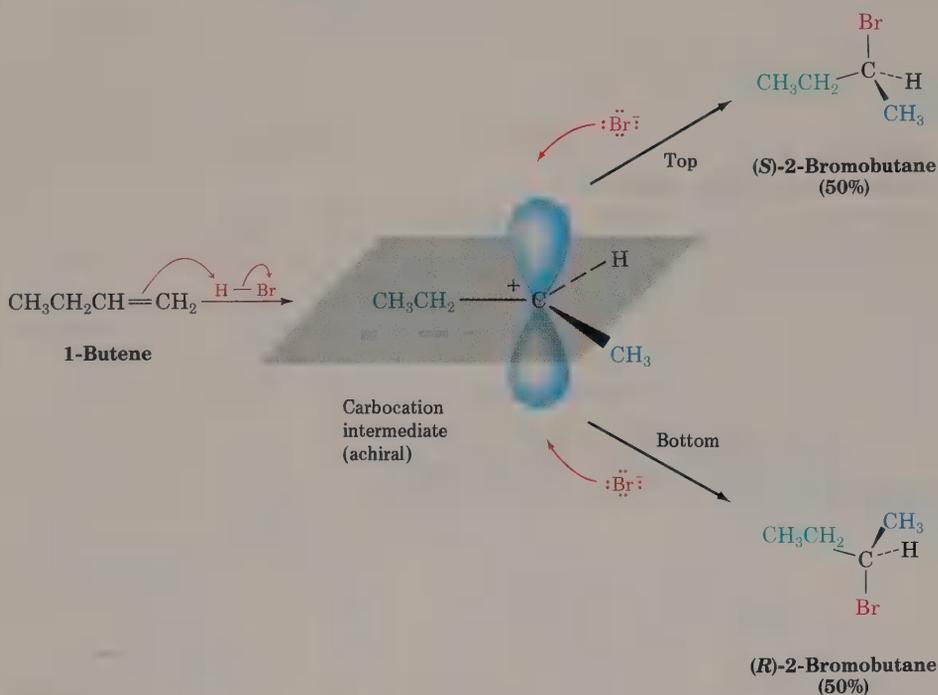


## 9.15 Stereochemistry of Reactions: Addition of HBr to Alkenes

Many organic reactions, including some that we've studied, yield products with stereogenic centers. For example, HBr adds to 1-butene to yield 2-bromobutane, a chiral molecule. What predictions can we make about the stereochemistry of this chiral product? If a single enantiomer is formed, is it  $R$  or  $S$ ? If a mixture of enantiomers is formed, how much of each is present? In fact, the 2-bromobutane produced is a racemic mixture of  $R$  and  $S$  enantiomers.

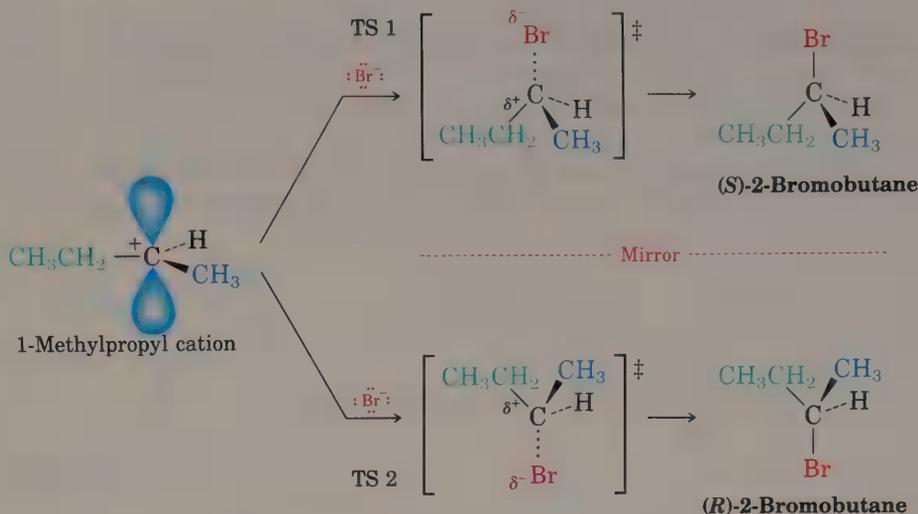


To understand *why* a racemic product results from reaction of HBr with 1-butene, think about what happens during the reaction. 1-Butene is first protonated by acid to yield an intermediate secondary ( $2^\circ$ ) carbocation. Since the trivalent carbon is  $sp^2$  hybridized, it has a plane of symmetry and is achiral. As a result, it can be attacked by bromide ion (also achiral) equally well from either the top or the bottom. Attack from the top leads to (*S*)-2-bromobutane, and attack from the bottom leads to (*R*)-2-bromobutane. Since both modes of attack are equally likely, a racemic mixture results (Figure 9.13).



**Figure 9.13** Stereochemistry of the addition of HBr to 1-butene. The intermediate achiral *sec*-butyl carbocation is attacked equally well from both top and bottom, leading to a racemic product mixture.

Another way to think about the reaction is in terms of transition states. If the intermediate carbocation is attacked from the top, *S* product is formed through transition state 1 (TS 1) in Figure 9.14. If the cation is attacked from the bottom, *R* product is formed through TS 2. *The two transition states are mirror images.* They therefore have identical activation energies and are equally likely to occur. Both transition states form at identical rates, leading to a 50:50 mixture of *R* and *S* products.



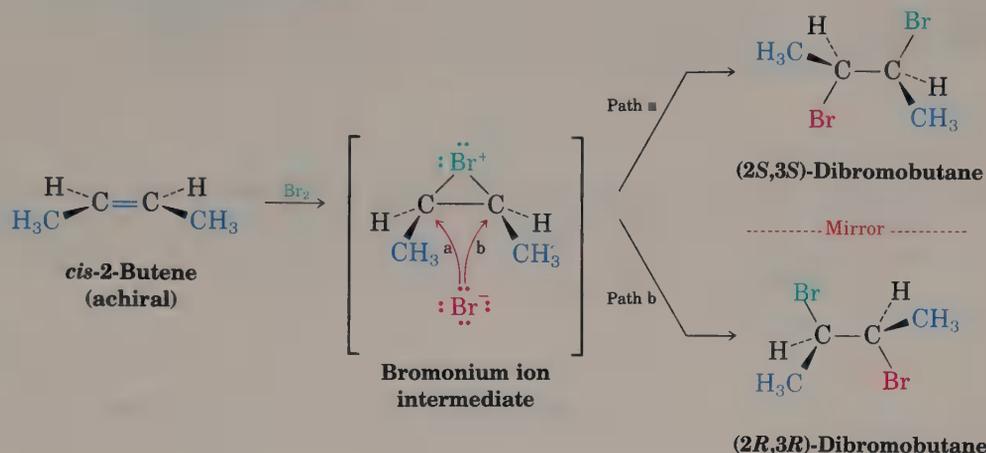
**Figure 9.14** Attack of bromide ion on the *sec*-butyl carbocation. Attack from the top leading to *S* product is the mirror image of attack from the bottom leading to *R* product. Since both are equally likely, racemic product is formed. The dotted C $\cdots$ Br bond in the transition state indicates partial bond formation.

## 9.16 Stereochemistry of Reactions: Addition of Br<sub>2</sub> to Alkenes

Addition of Br<sub>2</sub> to 2-butene leads to the formation of 2,3-dibromobutane and the generation of two stereogenic centers. What stereochemistry would we predict for such a reaction? If we begin with planar, achiral *cis*-2-butene, we would expect bromine to add to the double bond equally well from either the top or the bottom face to generate two intermediate bromonium ions. For the sake of simplicity, let's consider only the result of attack from the top face (Figure 9.15), keeping in mind that every structure we consider also has a mirror image.

The bromonium ion formed by top-face reaction of *cis*-2-butene can be attacked by bromide ion from either the right or the left side of the bottom face. Attack from the left (path a) leads to (2*S*,3*S*)-dibromobutane; attack from the right (path b) leads to (2*R*,3*R*)-dibromobutane. Since both modes of attack on the symmetrical, achiral bromonium ion are equally likely, a 50:50 (racemic) mixture of the two enantiomeric products is formed. Thus, we obtain (±)-2,3-dibromobutane.

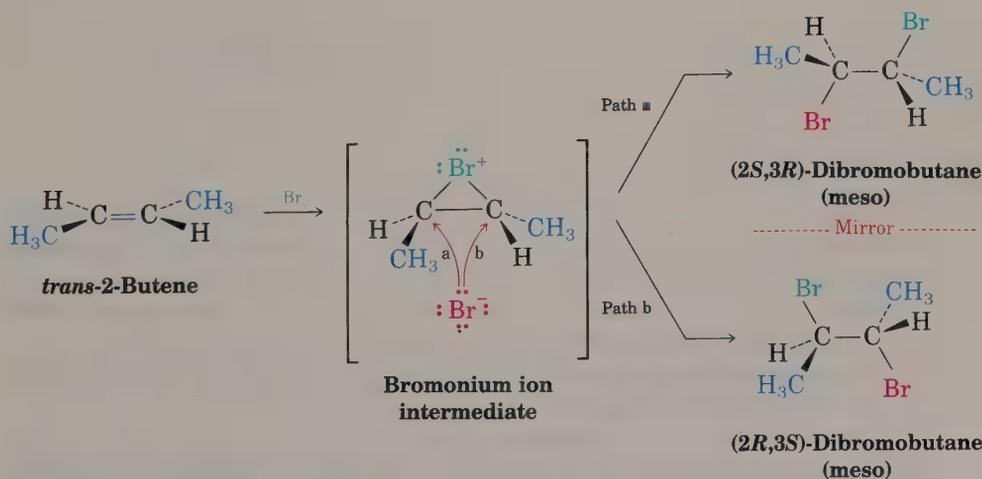
What about the addition of Br<sub>2</sub> to *trans*-2-butene? Is the same racemic product mixture formed? Perhaps surprisingly at first glance, the answer is no. *trans*-2-Butene reacts with bromine to form a bromonium ion, and once again we'll consider only top-face attack for simplicity (Figure 9.16). Attack of bromide ion on the bromonium ion intermediate takes place



**Figure 9.15** Stereochemistry of the addition of Br<sub>2</sub> to *cis*-2-butene. A racemic mixture of *2S,3S* and *2R,3R* products is formed.

equally well from both right and left sides of the bottom face, leading to the formation of *2R,3S* and *2S,3R* products in equal amounts. A close look at the two products, however, shows that they are *identical*: Both structures represent *meso*-2,3-dibromobutane.

The key conclusion from all three of the addition reactions just discussed is that an optically inactive product has been formed. This is a general rule: *Any reaction between two achiral partners always leads to optically inactive products—either racemic or meso.* Put another way, optical activity can't come from out of nowhere; optically active products can't be produced from optically inactive starting materials.



**Figure 9.16** Stereochemistry of the addition of Br<sub>2</sub> to *trans*-2-butene. A *meso* product is formed.

PROBLEM.....

- 9.17 What is the stereochemistry of the product that results from addition of Br<sub>2</sub> to cyclohexene? Is the product optically active? Explain.

PROBLEM.....

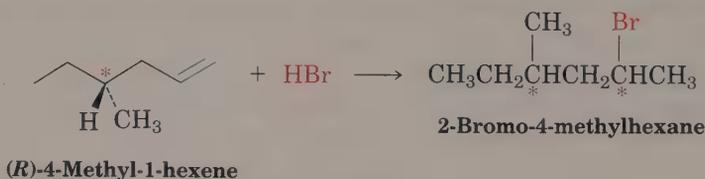
- 9.18 Addition of Br<sub>2</sub> to an unsymmetrical alkene such as *cis*-2-hexene leads to racemic product, even though attack of bromide ion on the unsymmetrical bromonium ion intermediate is not equally likely at both ends. Make drawings of the intermediate and the products, and explain the observed stereochemical result.

PROBLEM.....

- 9.19 Predict the stereochemical outcome of the reaction of Br<sub>2</sub> with *trans*-2-hexene. Show your reasoning.

## 9.17 Stereochemistry of Reactions: Addition of HBr to a Chiral Alkene

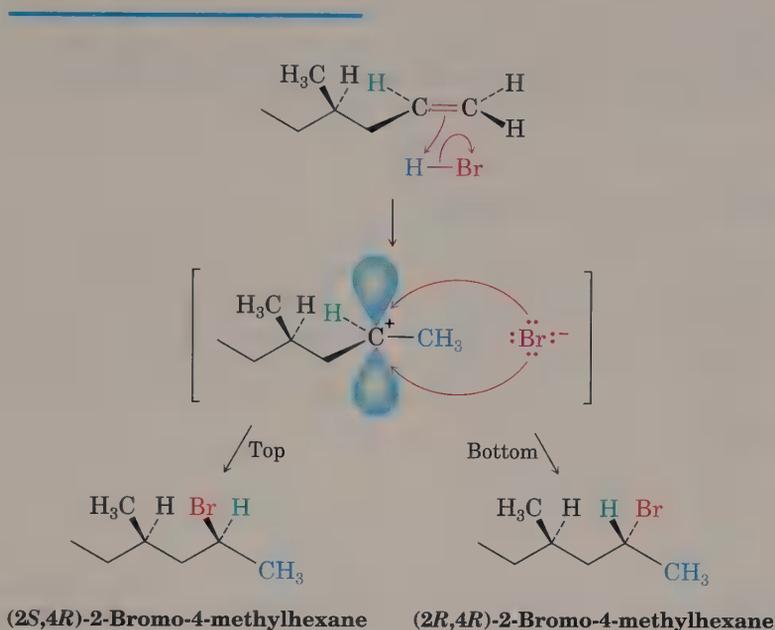
Both of the reactions considered in the two previous sections have involved additions to achiral alkene starting materials, and in both cases an optically inactive product was formed. What would happen, though, if we were to carry out a reaction on a single enantiomer of a *chiral* starting material? For example, what stereochemical result would be obtained from addition of HBr to a chiral alkene such as (*R*)-4-methyl-1-hexene? The product of this reaction, 2-bromo-4-methylhexane, has two stereogenic centers and might exist in any of four stereoisomeric configurations.



Let's think about the two stereogenic centers separately. What about the configuration at C4, the methyl-bearing carbon atom? Since C4 has the *R* configuration in the starting material, and since this stereogenic center is unaffected by the reaction, its configuration remains unchanged. Thus, the configuration of C4 in the product is also *R*.

What about the configuration at C2, the newly formed stereogenic center? As illustrated in Figure 9.17, the stereochemistry at C2 is established by bromide ion attack on a carbocation intermediate in the normal manner. *But this carbocation is not symmetrical; it is chiral because of the presence of the C4 center.* Since this intermediate carbocation has no plane of symmetry, it is not attacked equally well from top and bottom faces. One of the two faces is likely, for steric reasons, to be a bit more accessible than the

other face, leading to a mixture of *R* and *S* product stereochemistries in some ratio other than 50:50. Thus, the two products, (2*R*,4*R*)-2-bromo-4-methylhexane and (2*S*,4*R*)-2-bromo-4-methylhexane, are formed in unequal amounts.



**Figure 9.17** Stereochemistry of the addition of HBr to the chiral alkene, (*R*)-4-methyl-1-hexene. A mixture of diastereomeric 2*R*,4*R* and 2*S*,4*R* products is formed in a non-50:50 ratio because attack on the unsymmetrical intermediate carbocation is not equally likely from both top and bottom.

The two products formed in the reaction of achiral HBr with chiral alkene are diastereomers, and both are optically active. This is a general rule: *Any reaction between an achiral reagent and a chiral reagent leads to unequal amounts of diastereomeric products.* If the chiral starting material is optically active because only one enantiomer is used, then the products are also optically active.

PROBLEM.....

- 9.20 What products are formed and in what amounts from reaction of HBr with racemic ( $\pm$ )-4-methyl-1-hexene? Is the product mixture optically active?

PROBLEM.....

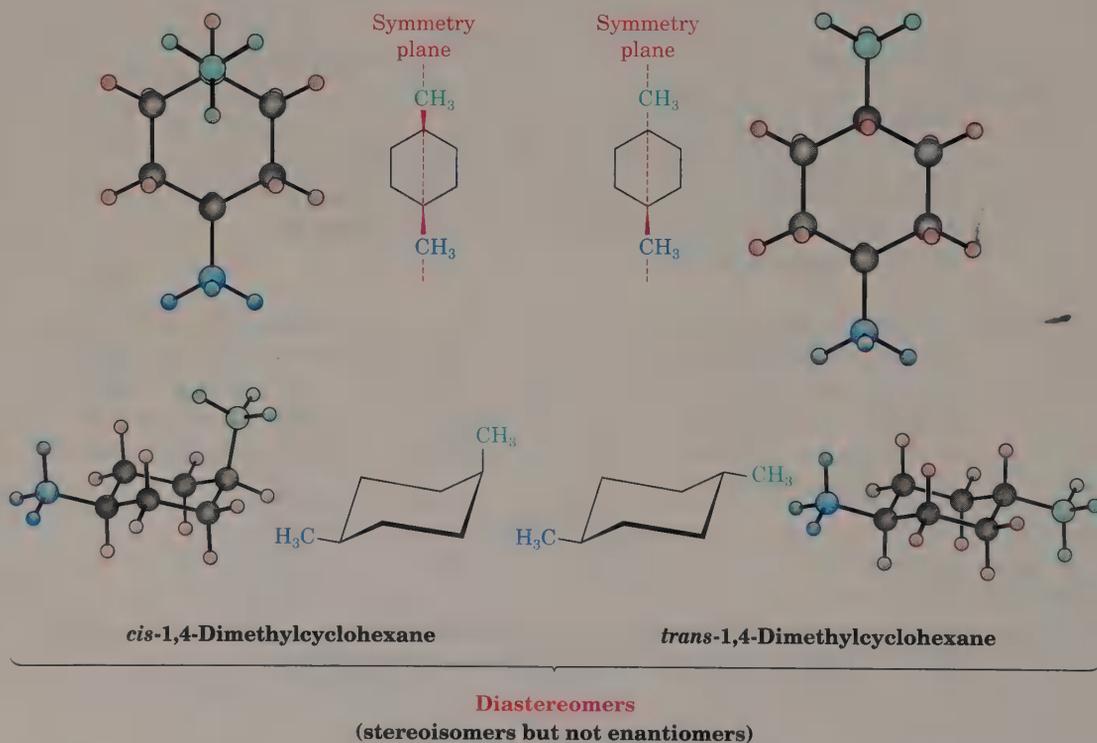
- 9.21 What products are formed and in what amounts from reaction of HBr with 4-methylcyclopentene?
- .....

## 9.18 Stereoisomerism and Chirality in Substituted Cyclohexanes

We saw in Section 4.13 that substituted cyclohexane rings adopt a chair-like geometry and that the conformation of a specific compound can be predicted by looking at steric interactions in the molecule. To complete a study of cyclohexane stereochemistry, we now need to examine the effect of conformation on stereoisomerism and chirality.

### 1,4-DISUBSTITUTED CYCLOHEXANES

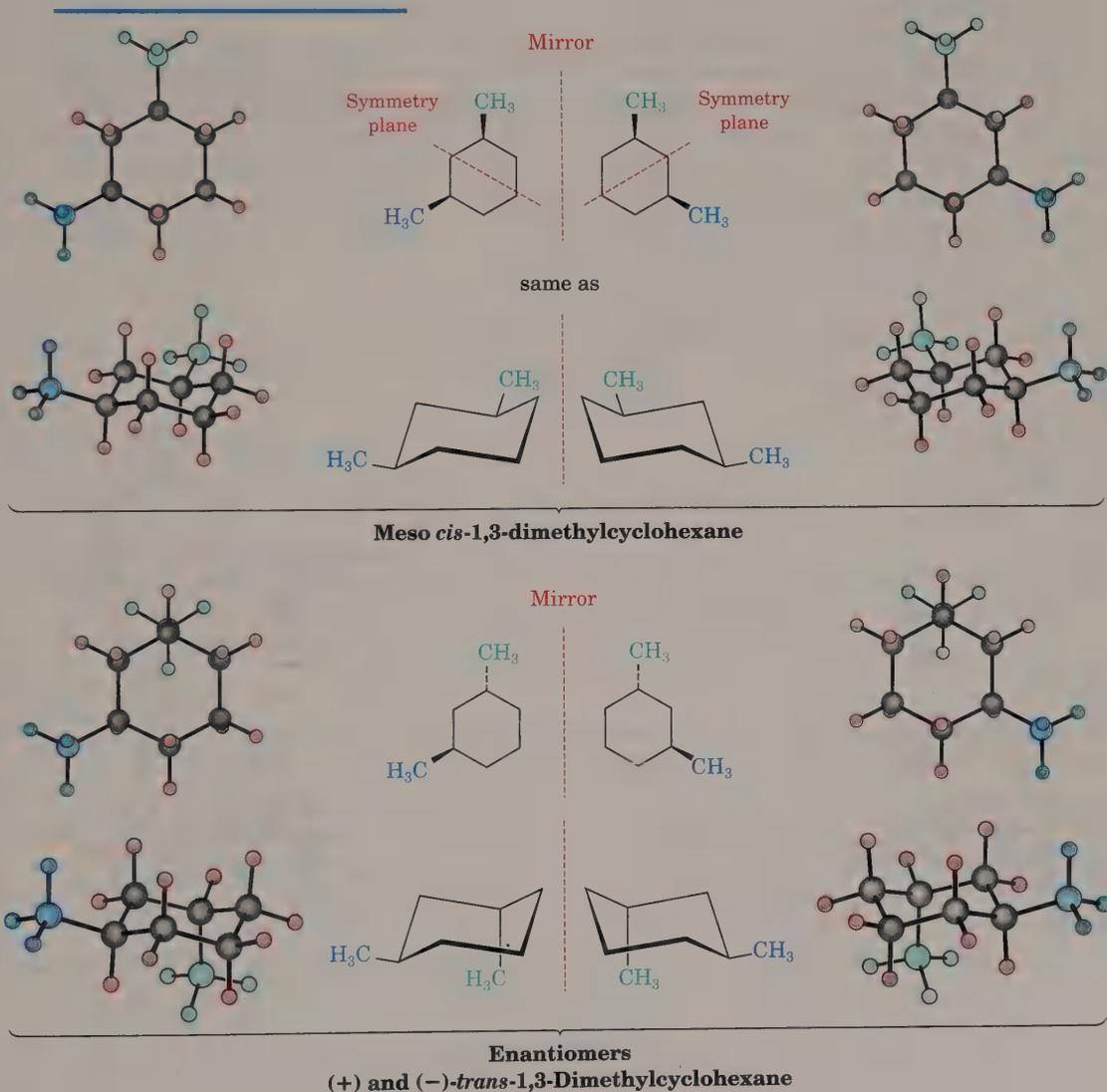
1,4-Disubstituted cyclohexanes have no stereogenic centers by virtue of a symmetry plane passing through the substituents and through carbons 1 and 4 of the ring. Thus, only cis and trans stereoisomers (diastereomers) are possible. The symmetry plane is evident whether we use a flat view of the molecule or a three-dimensional view of the chair conformation (Figure 9.18).



**Figure 9.18** The stereochemical relationships among 1,4-dimethylcyclohexanes. Both cis and trans isomers are achiral.

## 1,3-DISUBSTITUTED CYCLOHEXANES

1,3-Disubstituted cyclohexanes have two stereogenic centers, and a maximum of four stereoisomers is therefore possible. *cis*-1,3-Dimethylcyclohexane, however, has a symmetry plane and is thus a meso compound, whereas *trans*-1,3-dimethylcyclohexane has no symmetry plane and exists as a pair of enantiomers. Again, these symmetry properties are evident both from flat views and from conformational views (Figure 9.19).



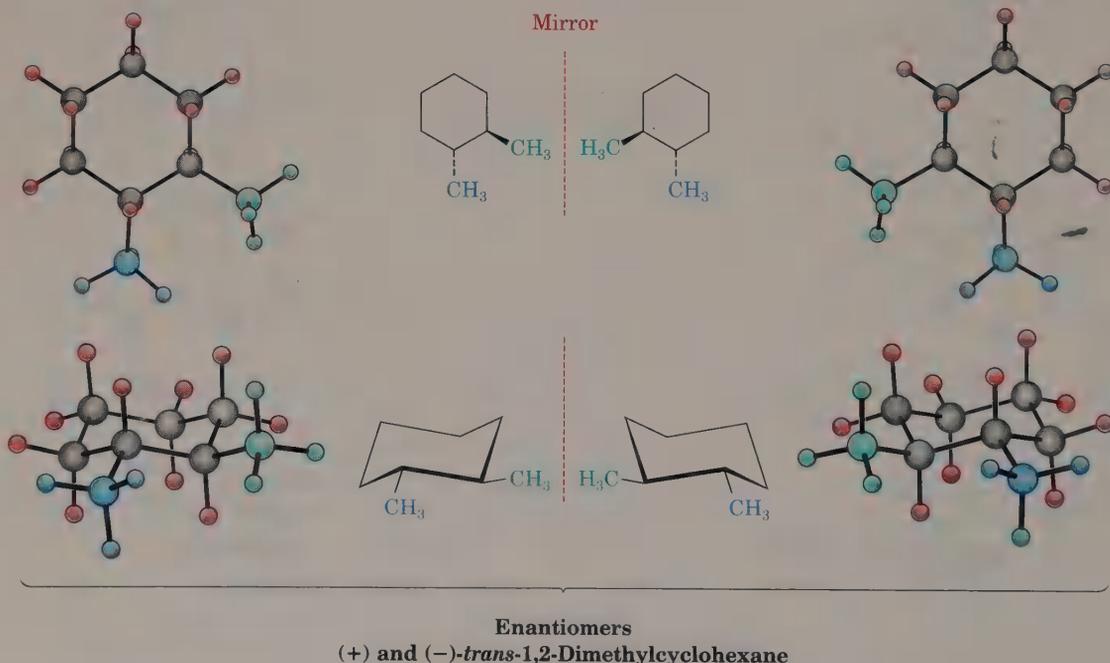
**Figure 9.19** The stereochemical relationships among 1,3-dimethylcyclohexanes. The *cis* isomer is a meso compound, and the *trans* isomer is a pair of enantiomers.

## 1,2-DISUBSTITUTED CYCLOHEXANES

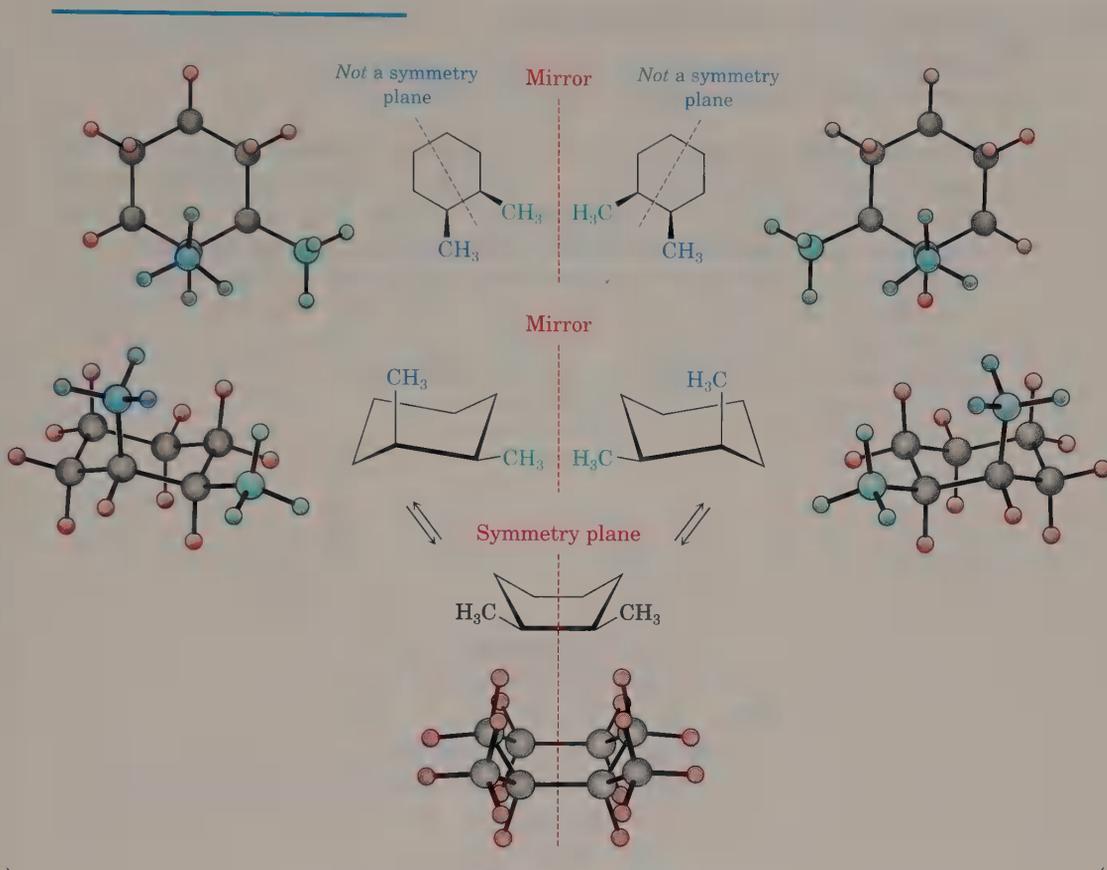
1,2-Disubstituted cyclohexanes have two stereogenic centers, and four stereoisomers are again possible. The *trans* isomer is straightforward: A top view of *trans*-1,2-dimethylcyclohexane shows that the molecule has no symmetry plane and is therefore a pair of (+) and (-) enantiomers (Figure 9.20).

The situation with the *cis* isomer is a bit more complicated than it appears at first glance. A flat view of *cis*-1,2-dimethylcyclohexane shows an apparent symmetry plane and leads to the conclusion that the *cis* isomer is an optically inactive *meso* compound. If the *cis* isomer is viewed in chair conformation, though, the symmetry plane is no longer present because of the puckering of the ring. Thus, we might now predict that *cis*-1,2-dimethylcyclohexane exists as a pair of (+) and (-) enantiomers.

In fact, *cis*-1,2-dimethylcyclohexane is a *meso* compound, just as predicted by the flat view. When flipped from the chair conformation to the boat conformation, it's apparent that both "enantiomers" give the same boat, which has a symmetry plane and is therefore *meso* (Figure 9.21). In order to be *meso*, it's not necessary that a compound has a symmetry plane in its most stable conformation; it's only necessary that the compound has a symmetry plane in *some* accessible conformation. Thus, we observe no optical activity for *cis*-1,2-dimethylcyclohexane.



**Figure 9.20** *trans*-1,2-Dimethylcyclohexane exists as a pair of (+) and (-) enantiomers.



*cis*-1,2-Dimethylcyclohexane, a *meso* compound

**Figure 9.21** *cis*-1,2-Dimethylcyclohexane appears to be a pair of enantiomers when viewed in the chair conformation but is actually a *meso* compound, as can be seen in the boat conformation.

The overall conclusion from examining all six dimethylcyclohexanes is that it is possible to predict the presence or absence of optical activity in any substituted cycloalkane merely by looking at the flat structures, without considering the exact three-dimensional conformation.

**PROBLEM** .....

- 9.22** How many stereoisomers of 1-chloro-3,5-dimethylcyclohexane are there? Draw the most stable conformation.

**PROBLEM** .....

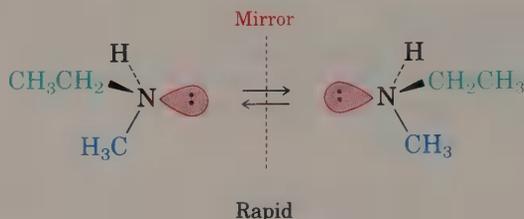
- 9.23** How many 1,2-dimethylcyclopentane stereoisomers are there? What are the stereochemical relationships among them?
- .....

## 9.19 Chirality at Atoms Other Than Carbon

Since the most common cause of chirality is the presence of four different substituents bonded to a tetrahedral atom, it follows that tetrahedral atoms other than carbon can also be stereogenic centers. Silicon, nitrogen, phosphorus, and sulfur are all commonly encountered in organic molecules, and all can be stereogenic centers under the proper circumstances. We know, for example, that trivalent nitrogen is tetrahedral and contains a lone pair of electrons (Section 1.12). Is trivalent nitrogen chiral? Does a compound such as ethylmethylamine exist as a pair of enantiomers?



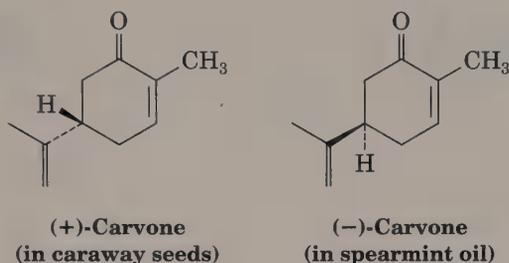
The answer is both yes and no. Yes in principle, but no in practice. Tetrahedral trivalent nitrogen compounds undergo a rapid “umbrella-like” inversion that interconverts enantiomers. We therefore can’t separate or isolate enantiomers except in special cases, and their existence is rarely important.



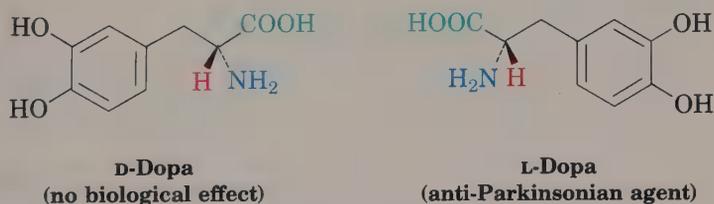
## 9.20 Chirality in Nature

We saw in Section 9.11 that the different stereoisomeric forms of tartaric acid have different physical properties. It’s usually the case that stereoisomers have different chemical and biological properties as well. For exam-

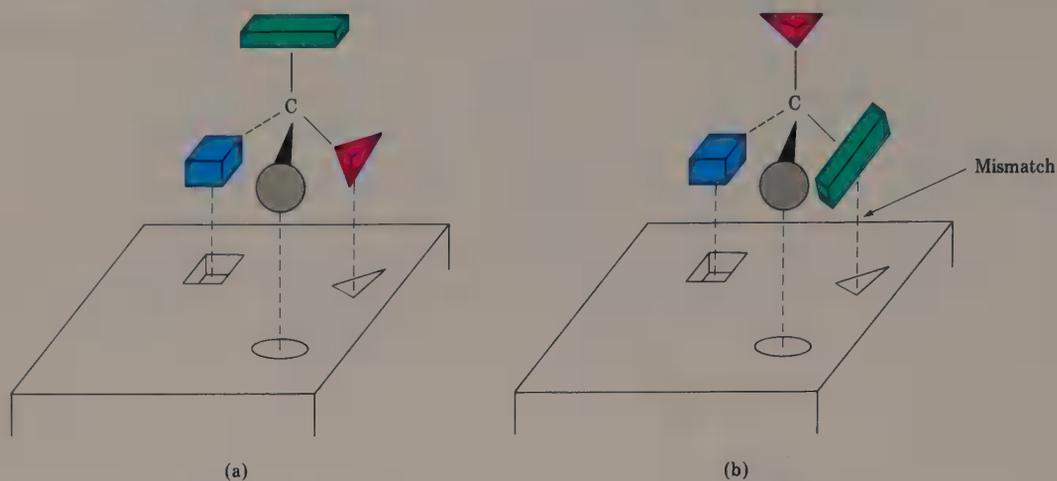
ple, the dextrorotatory enantiomer of carvone has the odor of caraway seeds, while the levorotatory enantiomer has the odor of spearmint.



Another particularly dramatic example of how a simple change in chirality can affect the biological properties of a molecule is found in the amino acid, dopa. Dopa, more properly named 2-amino-3-(3,4-dihydroxyphenyl)propanoic acid, has a single stereogenic center and can thus exist in two stereoisomeric forms. Although the dextrorotatory enantiomer, D-dopa, has no physiological effect on humans, the levorotatory enantiomer, L-dopa, is widely used for its potent activity against Parkinson's disease, a chronic malady of the central nervous system.



Why do different stereoisomers have such dramatically different biological properties? In order to exert its biological action, a chiral molecule must fit into a chiral receptor at the target site, much as a hand fits into a glove. Just as a right hand can fit into only a right glove, so a particular stereoisomer can fit into only a receptor having the proper complementary shape. Any other stereoisomer will be a misfit like a right hand in a left glove. A schematic representation of the interaction between a chiral molecule and a chiral biological receptor is shown in Figure 9.22. One enantiomer fits the receptor perfectly, but the other does not.



**Figure 9.22** (a) One enantiomer fits easily into a chiral receptor site to exert its biological effect, but (b) the other enantiomer can't fit into the same receptor and is thus without biological effect.

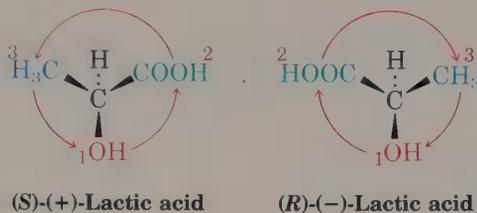
## 9.21 Summary and Key Words

When a beam of **plane-polarized light** is passed through a solution of certain organic molecules, the plane of polarization is rotated. Compounds that exhibit this behavior are called **optically active**. Optical activity is due to the asymmetric structure of the molecules themselves.

An object or molecule that is not superimposable on its mirror image is said to be **chiral**, meaning “handed.” For example, a glove is chiral but a coffee cup is nonchiral, or **achiral**. A chiral molecule is one that does not contain a **plane of symmetry**—an imaginary plane that cuts through the molecule so that one half is a mirror image of the other half. The most common cause of chirality in organic molecules is the presence of a tetrahedral,  $sp^3$ -hybridized carbon atom bonded to four different groups. Compounds that contain such chiral carbon atoms exist as a pair of non-superimposable, mirror-image stereoisomers called **enantiomers**. Enantiomers are identical in all physical properties except for the direction in which they rotate plane-polarized light.

The stereochemical **configuration** of a chiral carbon atom can be depicted using **Fischer projections**, in which horizontal lines (bonds) are understood to come out of the plane of the paper and vertical bonds are understood to go back into the plane of the paper. The configuration can be specified as either *R* (*rectus*) or *S* (*sinister*) by using the Cahn–Ingold–Prelog sequence rules. This is done by first assigning priorities to the four substituents on the chiral carbon atom and then orienting the molecule so that

the lowest-priority group points directly back away from the viewer. We then look at the remaining three substituents and let the eye travel from the group having the highest priority to second highest to third highest. If the direction of travel is clockwise, the configuration is labeled *R*; if the direction of travel is counterclockwise, the configuration is labeled *S*.



Some molecules have more than one stereogenic center. Enantiomers have opposite configuration at all stereogenic centers, whereas **diastereomers** have the same configuration in at least one center but opposite configurations at the others. A compound with  $n$  stereogenic centers can have  $2^n$  stereoisomers.

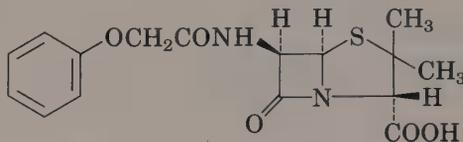
**Meso compounds** contain stereogenic centers, but are achiral overall because they contain a plane of symmetry. **Racemic mixtures**, or **racemates**, are 50:50 mixtures of (+) and (–) enantiomers. Racemic mixtures and individual diastereomers differ from each other in their physical properties such as solubility, melting point, and boiling point.

Most reactions give chiral products. If the starting materials are optically inactive, the products must also be optically inactive—either meso or racemic. If one or both of the starting materials is optically active, however, the product will also be optically active if the original stereogenic center remains.

## ADDITIONAL PROBLEMS .....

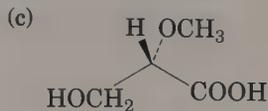
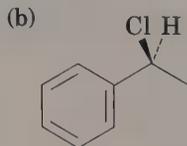
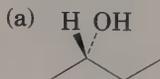
- 9.24 Cholic acid, the major steroid found in bile, was observed to have a specific rotation of  $+2.22^\circ$  when a 3.0 gram sample was dissolved in 5.0 mL alcohol and the solution was placed in a sample tube with a 1.0 cm path length. Calculate  $[\alpha]_D$  for cholic acid.
- 9.25 Polarimeters for measuring optical rotation are quite sensitive and can measure rotations to 0.001 degree, an important fact when only small amounts of sample are available. Ecdysone, for example, is an insect hormone that controls molting in the silkworm moth. When 7.0 mg ecdysone was dissolved in 1.0 mL chloroform and the solution was placed in a 2.0 cm path-length cell, an observed rotation of  $+0.087^\circ$  was found. Calculate  $[\alpha]_D$  for ecdysone.
- 9.26 What do these terms mean?
- |                  |                        |                      |
|------------------|------------------------|----------------------|
| (a) Chirality    | (b) Stereogenic center | (c) Optical activity |
| (d) Diastereomer | (e) Enantiomer         | (f) Racemate         |
- 9.27 Which of these compounds are chiral? Draw them and label the stereogenic centers.
- |   |                                 |
|---|---------------------------------|
| (a) 2,4-Dimethylheptane                 | (b) 3-Ethyl-5,5-dimethylheptane |
| (c) <i>cis</i> -1,4-Dichlorocyclohexane | (d) 4,5-Dimethyl-2,6-octadiyne  |

- 9.28 Draw chiral molecules that meet these descriptions:  
 (a) A chloroalkane,  $C_5H_{11}Cl$  (b) An alcohol,  $C_6H_{14}O$   
 (c) An alkene,  $C_6H_{12}$  (d) An alkane,  $C_8H_{18}$
- 9.29 Eight alcohols have the formula  $C_5H_{12}O$ . Draw them. Which are chiral?
- 9.30 Draw the nine chiral molecules that have the formula  $C_6H_{13}Br$ .
- 9.31 Draw compounds that fit these descriptions:  
 (a) A chiral alcohol with four carbons  
 (b) A chiral carboxylic acid having the formula  $C_5H_{10}O_2$   
 (c) A compound with two stereogenic centers  
 (d) A chiral aldehyde having the formula  $C_3H_5BrO$
- 9.32 Which of these objects are chiral?  
 (a) A basketball (b) A fork  
 (c) A wine glass (d) A golf club  
 (e) A monkey wrench (f) A snowflake
- 9.33 Penicillin V is an important broad-spectrum antibiotic that contains three stereogenic centers. Identify them with asterisks.

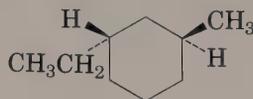
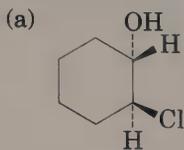


Penicillin V  
(antibiotic)

- 9.34 Draw examples of the following:  
 (a) A meso compound with the formula  $C_8H_{18}$   
 (b) A meso compound with the formula  $C_9H_{20}$   
 (c) A compound with two stereogenic centers, one *R* and the other *S*
- 9.35 What is the relationship between the specific rotations of (2*R*,3*R*)-dichloropentane and (2*S*,3*S*)-dichloropentane? Between (2*R*,3*S*)-dichloropentane and (2*R*,3*R*)-dichloropentane?
- 9.36 What is the stereochemical configuration of the enantiomer of (2*S*,4*R*)-dibromooctane?
- 9.37 What are the stereochemical configurations of the two diastereomers of (2*S*,4*R*)-dibromooctane?
- 9.38 Assign Cahn-Ingold-Prelog priorities to these sets of substituents:  
 (a)  $-CH=CH_2$ ,  $-CH(CH_3)_2$ ,  $-C(CH_3)_3$ ,  $-CH_2CH_3$   
 (b)  $-C\equiv CH$ ,  $-CH=CH_2$ ,  $-C(CH_3)_3$ ,   
 (c)  $-CO_2CH_3$ ,  $-COCH_3$ ,  $-CH_2OCH_3$ ,  $-CH_2CH_3$   
 (d)  $-C\equiv N$ ,  $-CH_2Br$ ,  $-CH_2CH_2Br$ ,  $-Br$
- 9.39 Assign *R,S* configurations to the stereogenic centers in these molecules:



9.40 Assign *R* or *S* configurations to each stereogenic center in these molecules:

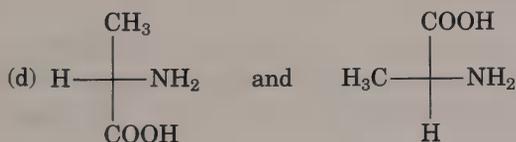
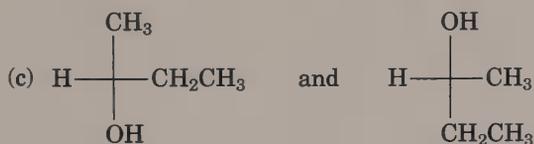
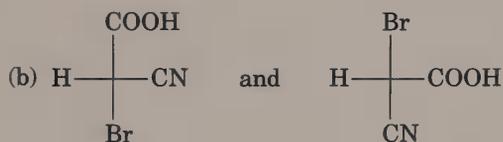
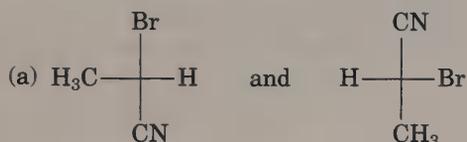


9.41 Draw tetrahedral representations of these molecules:

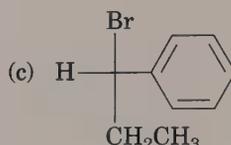
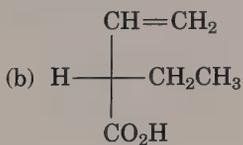
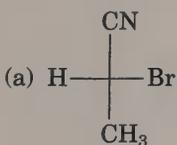
(a) (*S*)-2-Butanol,  $\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_3$  (b) (*R*)-3-Chloro-1-pentene

9.42 Draw tetrahedral representations of the two enantiomers of the amino acid cysteine,  $\text{HSCH}_2\text{CH}(\text{NH}_2)\text{COOH}$ , and identify each as *R* or *S*.

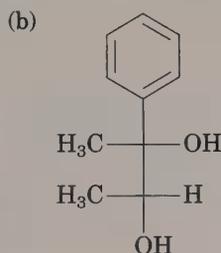
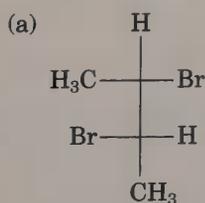
9.43 Which of these pairs of Fischer projections represent the same enantiomer, and which represent different enantiomers?

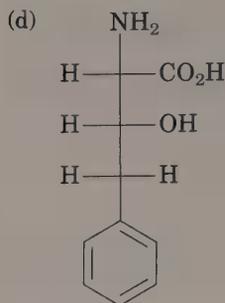
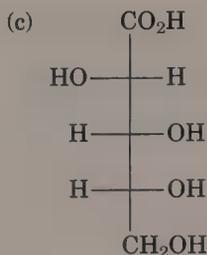


9.44 Assign *R,S* configurations to these Fischer projections:



9.45 Assign *R* or *S* configurations to each stereogenic center in these molecules:

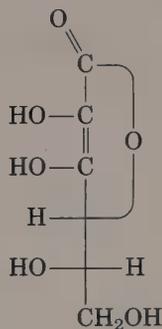




9.46 Draw Fischer projections that fit these descriptions:

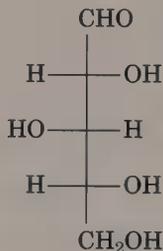
- The *S* enantiomer of 2-bromobutane
- The *R* enantiomer of alanine,  $\text{CH}_3\text{CH}(\text{NH}_2)\text{COOH}$
- The *R* enantiomer of 2-hydroxypropanoic acid
- The *S* enantiomer of 3-methylhexane

9.47 Assign *R* or *S* configurations to each chiral carbon atom in ascorbic acid (vitamin C).



Ascorbic acid

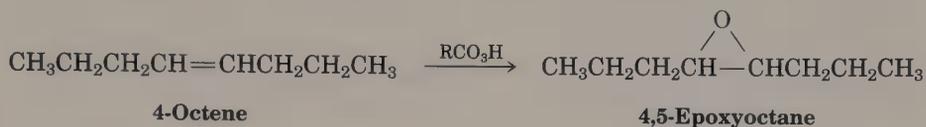
9.48 Xylose is a common sugar found in many woods, including maple and cherry. Because it is much less prone to cause tooth decay than sucrose, xylose has been used in candy and chewing gum. Assign *R,S* configurations to the stereogenic centers in xylose:



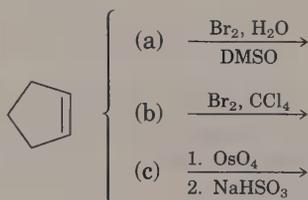
(+)-Xylose,  $[\alpha]_{\text{D}} = +92^\circ$

- 9.49 Hydroxylation of *cis*-2-butene with  $\text{OsO}_4$  yields 2,3-butanediol. What stereochemistry do you expect for the product? (Review Section 7.8 if necessary.)
- 9.50 Hydroxylation of *trans*-2-butene with  $\text{OsO}_4$  also yields 2,3-butanediol. What stereochemistry do you expect for the product?

- 9.51 Alkenes undergo reaction with peroxycarboxylic acids ( $\text{RCO}_3\text{H}$ ) to give three-membered-ring cyclic ethers called *epoxides*. For example, 4-octene reacts with peroxyacids to yield 4,5-epoxyoctane:

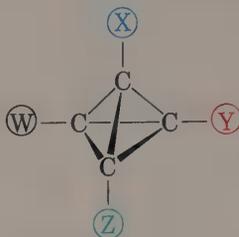


- Assuming that this epoxidation reaction occurs with *syn* stereochemistry, draw the structure obtained from epoxidation of *cis*-4-octene. Is this epoxide chiral? How many stereogenic centers does it have? How would you describe the product stereochemically?
- 9.52 Answer Problem 9.51, assuming that the epoxidation reaction is carried out on *trans*-4-octene.
- 9.53 Write the products of the following reactions, and indicate the stereochemistry obtained in each instance.



- 9.54 Draw all possible stereoisomers of cyclobutane-1,2-dicarboxylic acid, and indicate the interrelationships. Which, if any, are optically active? Do the same for cyclobutane-1,3-dicarboxylic acid.
- 9.55 Compound A,  $\text{C}_7\text{H}_{12}$ , was found to be optically active. On catalytic reduction over a palladium catalyst, two equivalents of hydrogen were absorbed, yielding compound B,  $\text{C}_7\text{H}_{16}$ . On ozonolysis, two fragments were obtained. One fragment was identified as acetic acid. The other fragment, compound C, was an optically active carboxylic acid,  $\text{C}_5\text{H}_{10}\text{O}_2$ . Formulate the reactions, and draw structures for A, B, and C.
- 9.56 Compound A,  $\text{C}_{11}\text{H}_{16}\text{O}$ , was found to be an optically active alcohol. Despite its apparent unsaturation, no hydrogen was absorbed on catalytic reduction over a palladium catalyst. On treatment of A with dilute sulfuric acid, dehydration occurred, and an optically inactive alkene B,  $\text{C}_{11}\text{H}_{14}$ , was produced as the major product. Alkene B, on ozonolysis, gave two products. One product was identified as propanal,  $\text{CH}_3\text{CH}_2\text{CHO}$ . Compound C, the other product,  $\text{C}_8\text{H}_8\text{O}$ , was shown to be a ketone. How many multiple bonds and/or rings does A have? Formulate the reactions and identify A, B, and C.
- 9.57 Draw the structure of (*R*)-2-methylcyclohexanone.
- 9.58 How many stereoisomers of 2,4-dibromo-3-chloropentane are there? Draw them and indicate which are optically active.
- 9.59 The so-called tetrahedranes are an interesting class of compounds, the first example of which was synthesized in 1979. Construct a model (carefully!) of tetrahedrane.

Consider a substituted tetrahedrane with four different substituents. Is it chiral? Explain your answer.

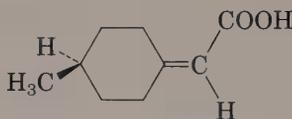


- 9.60** *Allenes*, compounds with adjacent carbon-carbon double bonds, are well known. Many allenes are chiral, even though they don't contain chiral carbon atoms. Mycomycin, for example, a naturally occurring antibiotic isolated from the bacterium *Nocardia acidophilus*, is chiral and has  $[\alpha]_D = -130^\circ$ . Why is mycomycin chiral? Making a molecular model should be helpful.



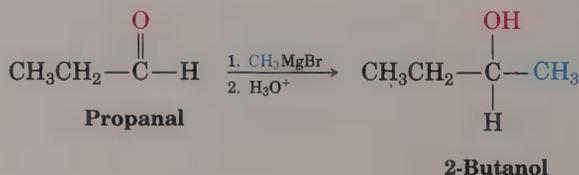
**Mycomycin**  
(an allene)

- 9.61** Long before optically active allenes were known (Problem 9.60), the resolution of 4-methylcyclohexylideneacetic acid into two enantiomers had been carried out. This was the first molecule to be resolved that was chiral yet did not contain a stereogenic center. Why is this molecule chiral? What geometric relation does this molecule have to allenes?

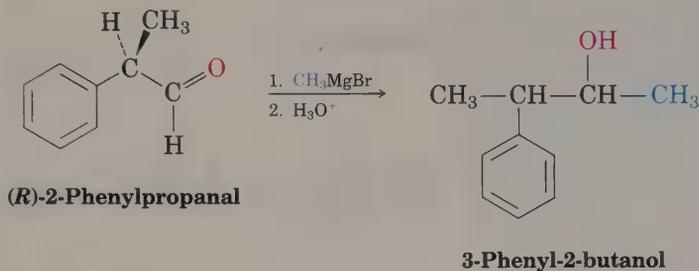


**4-Methylcyclohexylideneacetic acid**

- 9.62** (*S*)-1-Chloro-2-methylbutane undergoes light-induced reaction with  $\text{Cl}_2$  by a radical mechanism to yield a mixture of products. Among the products are 1,4-dichloro-2-methylbutane and 1,2-dichloro-2-methylbutane.
- Formulate the reaction, showing the correct stereochemistry of the starting material.
  - One of the two products is optically active, but the other is optically inactive. Which is which?
  - What can you conclude about the stereochemistry of radical chlorination reactions based on the results of this experiment?
- 9.63** Grignard reagents,  $\text{RMgX}$ , react with aldehydes to yield alcohols. For example, the reaction of methylmagnesium bromide with propanal yields 2-butanol:



- (a) Is the product chiral? Is it optically active?  
 (b) How many stereoisomers of butanol are formed, what are their stereochemical relationships, and what are their relative percentages?
- 9.64 Imagine that another Grignard reaction similar to that in Problem 9.63 is carried out between methylmagnesium bromide and (*R*)-2-phenylpropanal to yield 3-phenyl-2-butanol:



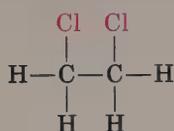
- (a) Is the product chiral? Is it optically active?  
 (b) How many stereoisomers of 3-phenyl-2-butanol are formed, what are their stereochemical relationships, and what are their relative percentages?

# Chapter 10

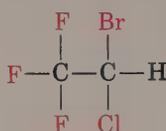


## Alkyl Halides

It would be difficult to study organic chemistry for long without becoming aware of halo-substituted alkanes. Among their many uses, alkyl halides are employed as industrial solvents, as inhaled anesthetics in medicine, as refrigerants, and as pesticides and fumigating agents.



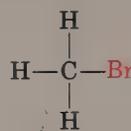
1,2-Dichloroethane  
(a solvent)



Halothane  
(an inhaled anesthetic)

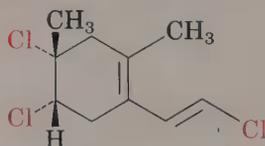


Freon 12  
(a refrigerant)



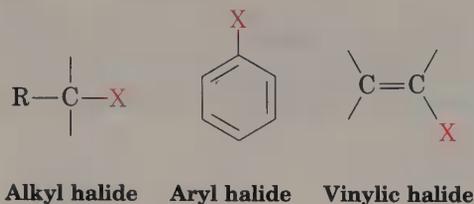
Bromomethane  
(a fumigant)

Alkyl halides also occur widely in nature, though mostly in marine rather than terrestrial organisms. The 1980s saw an explosive growth in the chemical investigation of marine organisms, and the structures of many naturally occurring halogenated molecules have been elucidated. Thus, simple halomethanes such as  $\text{CHCl}_3$ ,  $\text{CCl}_4$ ,  $\text{CBr}_4$ ,  $\text{CH}_3\text{I}$ , and  $\text{CH}_3\text{Cl}$  are constituents of the Hawaiian alga *Asparagopsis taxiformis*. In addition, many substances isolated from marine organisms exhibit interesting biological activity. For example, plocamene B, a trichlorocyclohexene derivative isolated from the red alga *Plocamium violaceum*, is similar in potency to DDT in showing insecticidal activity against mosquito larvae.



Plocamene B  
(a trichloride)

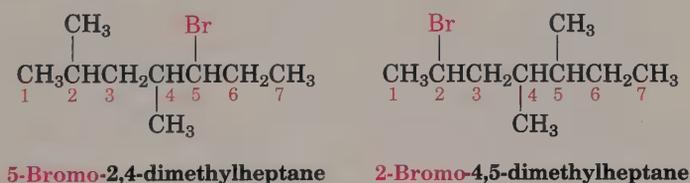
Before discussing the chemistry of alkyl halides, it should be pointed out that we will be talking primarily about compounds having halogen atoms bonded to saturated,  $sp^3$ -hybridized carbon. Other classes of organohalides, such as aromatic (*aryl*) and alkenyl (*vinyllic*) halides, also exist, but much of their chemistry is different.



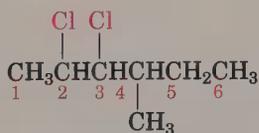
## 10.1 Nomenclature of Alkyl Halides

Alkyl halides are named in the same way that alkanes are (Section 3.4) by considering the halogen as a substituent on a parent alkane chain. There are three rules:

1. Find and name the parent chain. As in naming alkanes, select the longest chain as the parent. If a double or triple bond is present, the parent chain must contain it.
2. Number the carbons of the parent chain beginning at the end nearer the first substituent, regardless of whether it is alkyl or halo. Assign each substituent a number according to its position on the chain. For example:

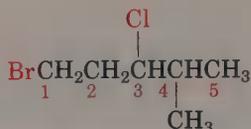


- a. If more than one of the same kind of halogen is present, number each and use one of the prefixes *di-*, *tri-*, *tetra-*, and so on. For example:

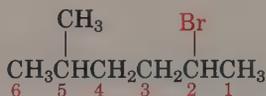


2,3-Dichloro-4-methylhexane

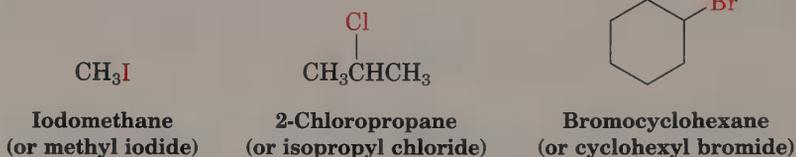
- b. If different halogens are present, number each according to its position on the chain, but list all substituents in alphabetical order when writing the name. For example:

**1-Bromo-3-chloro-4-methylpentane**

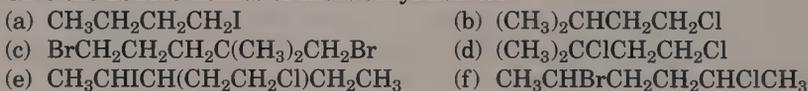
3. If the parent chain can be properly numbered from either end by rule 2, begin at the end nearer the substituent (either alkyl or halo) that has alphabetical precedence. For example:

**2-Bromo-5-methylhexane**  
(not 5-Bromo-2-methylhexane)

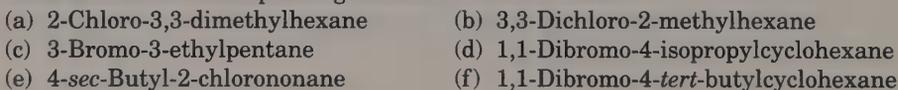
In addition to their systematic names, many simple alkyl halides are also named by identifying first the alkyl group and then the halogen. For example,  $\text{CH}_3\text{I}$  can be called either iodomethane or methyl iodide. Such names are well entrenched in the chemical literature and in daily usage, but they won't be used in this book.

**PROBLEM** .....

- 10.1**
- Give the IUPAC names of these alkyl halides:

**PROBLEM** .....

- 10.2**
- Draw structures corresponding to these IUPAC names:

**10.2 Structure of Alkyl Halides**

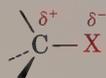
The carbon-halogen bond in an alkyl halide results from the overlap of a carbon  $sp^3$  hybrid orbital with a halogen orbital. Thus, alkyl halide carbon atoms have an approximately tetrahedral geometry with H-C-X bond

angles near  $109^\circ$ . Halogens increase in size going down the periodic table, an increase that is reflected in the bond lengths of the halomethane series (Table 10.1). Table 10.1 also indicates that C–X bond strengths decrease going down the periodic table.

**TABLE 10.1** Parameters for the C–X Bond in Halomethanes

<i>Halomethane</i>	<i>Bond length (Å)</i>	<i>Bond strength</i>	
		<i>(kcal/mol)</i>	<i>(kJ/mol)</i>
H <sub>3</sub> C—F	1.39	109	456
H <sub>3</sub> C—Cl	1.78	84	351
H <sub>3</sub> C—Br	1.93	70	293
H <sub>3</sub> C—I	2.14	56	234

In an earlier discussion of bond polarity in functional groups (Section 5.5), we noted that halogens are electronegative with respect to carbon. The C–X bond is polar, with the carbon atom bearing a slight positive charge ( $\delta^+$ ) and the halogen a slight negative charge ( $\delta^-$ ):



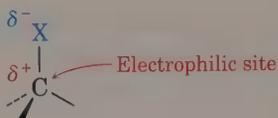
where X is the standard abbreviation for a halogen (X = F, Cl, Br, or I).

We can get a rough idea of the amount of bond polarity by measuring the dipole moments of the halomethanes. As indicated in Table 10.2, the identity of the halogen has a fairly small effect; all of the halomethanes have substantial dipole moments.

**TABLE 10.2** Dipole Moments of Halomethanes

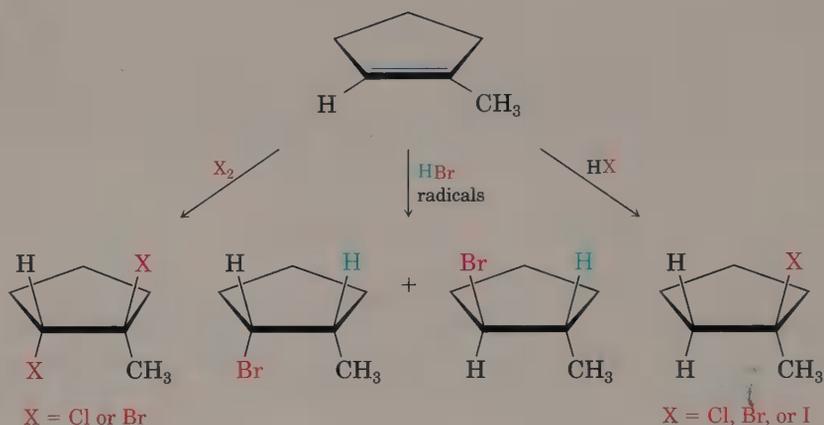
<i>Halomethane</i>	<i>Dipole moment, <math>\mu</math> (D)</i>
CH <sub>3</sub> F	1.82
CH <sub>3</sub> Cl	1.94
CH <sub>3</sub> Br	1.79
CH <sub>3</sub> I	1.64

Since the carbon atom of alkyl halides is positively polarized, alkyl halides are good electrophiles (Section 5.5). We'll see in the next chapter that much of the chemistry alkyl halides undergo is dominated by their electrophilic behavior.



### 10.3 Preparation of Alkyl Halides

We've already seen several methods of alkyl halide preparation, including the reaction of both  $\text{HX}$  and  $\text{X}_2$  with alkenes in electrophilic addition reactions (Sections 6.9 and 7.2).  $\text{HCl}$ ,  $\text{HBr}$ , and  $\text{HI}$  react with alkenes by a polar pathway to give the product of Markovnikov addition;  $\text{HBr}$  can also add by a radical pathway to give the non-Markovnikov product. Bromine and chlorine yield trans 1,2-dihalogenated addition products.



Another method of alkyl halide synthesis is the reaction of alkanes with chlorine or bromine by a radical chain-reaction pathway (Section 5.4). Although inert to most reagents, alkanes react readily with chlorine or bromine gas in the presence of light to give haloalkane substitution products. The reaction occurs by the radical mechanism shown in Figure 10.1 for chlorination.

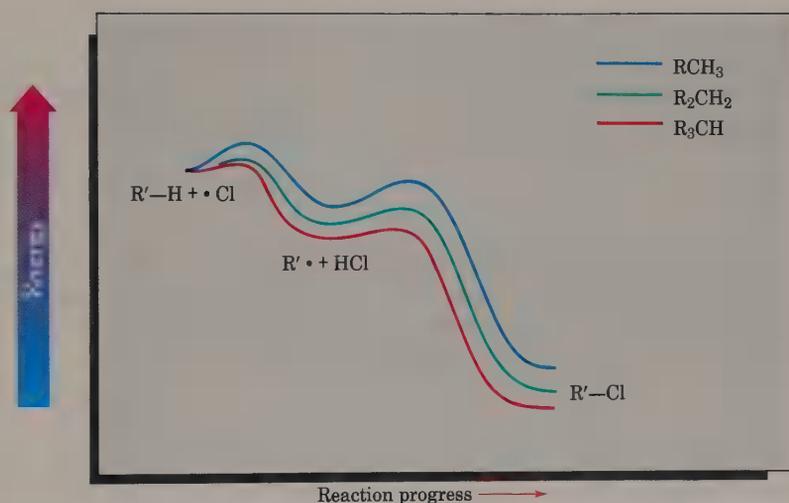
Recall from Section 5.4 that radical substitution reactions normally require three kinds of steps: initiation, propagation, and termination. Once an *initiation* step has started the process by producing radicals, the reaction continues in a self-sustaining cycle. The cycle requires two repeating *propagation* steps, in which a radical, the halogen, and the alkane yield alkyl halide product plus more radical to carry on the chain. The chain is occasionally *terminated* by the combination of two radicals.

Though interesting from a mechanistic point of view, alkane halogenation is a poor synthetic method for preparing different haloalkanes. Let's see why.



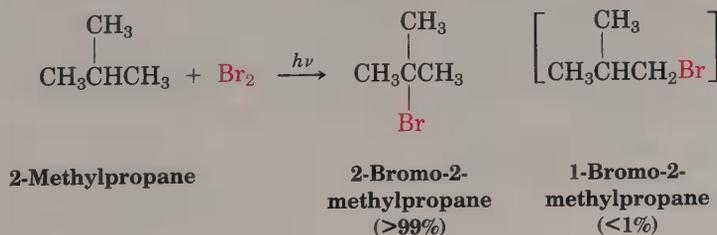


An explanation of the relationship between reactivity and bond strength in radical chlorination reactions is similar to that invoked in Section 6.12 to explain why more stable carbocations form faster than less stable ones in alkene electrophilic addition reactions. The reaction energy diagram for the formation of an alkyl radical during alkane chlorination looks like that in Figure 10.2. Although the hydrogen-abstraction step is slightly exothermic, there is nevertheless a certain amount of developing radical character in the transition state for this step. Thus, any factor (such as increased alkyl substitution) that stabilizes a radical intermediate also stabilizes the transition state leading to that intermediate (lowers its  $\Delta G^\ddagger$ ). In other words, a more stable radical forms faster than a less stable one. Tertiary radicals are more stable, and tertiary hydrogen atoms are therefore more easily removed.

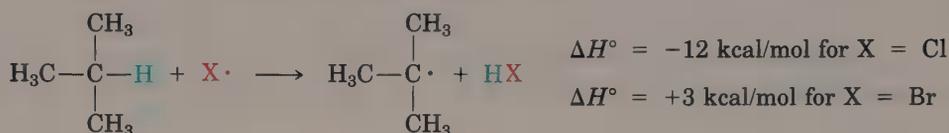


**Figure 10.2** Reaction energy diagram for alkane chlorination. The stability order of tertiary, secondary, and primary radicals is the same as their relative rate of formation.

In contrast to alkane chlorination, alkane bromination is usually a good deal more selective. In its reaction with 2-methylpropane, for example, bromine abstracts the tertiary hydrogen with greater than 99% selectivity, as opposed to the 35:65 mixture observed in the corresponding chlorination.



The enhanced selectivity of alkane bromination over chlorination can be explained by turning once again to the Hammond postulate. In comparing the product-determining steps—abstraction of an alkane hydrogen by  $\text{Cl}\cdot$  or  $\text{Br}\cdot$  radical—a comparison of  $\Delta H^\circ$  values shows that chlorinations are exothermic but that brominations are endothermic. As a result, the transition state for bromination resembles the alkyl radical more closely than the transition state for chlorination does, and the stability of that radical is therefore more important for bromination than for chlorination.



### 2-Methylpropane

#### PROBLEM.....

- 10.3 Draw and name all monochloro products you would expect to obtain from radical chlorination of 2-methylpentane. Which, if any, are chiral?

#### PROBLEM.....

- 10.4 Taking the known relative reactivities of  $1^\circ$ ,  $2^\circ$ , and  $3^\circ$  hydrogen atoms into account, what product(s) would you expect to obtain from monochlorination of 2-methylbutane? What would be the approximate percentages of each product? (Don't forget to take into account the number of each type of hydrogen.)

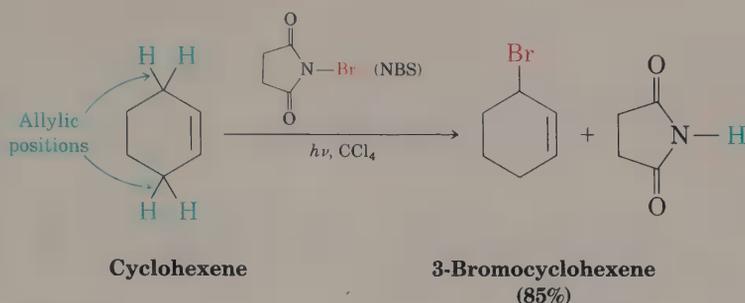
#### PROBLEM.....

- 10.5 Use the data in Table 5.4 to calculate  $\Delta H^\circ$  for the reactions of  $\text{Cl}\cdot$  and  $\text{Br}\cdot$  with a secondary hydrogen atom of propane. Which reaction would you expect to be more selective?
- .....

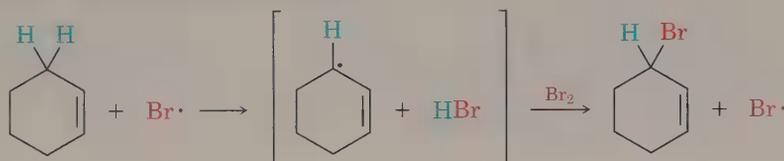
## 10.5 Allylic Bromination of Alkenes

While repeating some earlier work of Wohl,<sup>1</sup> Karl Ziegler reported in 1942 that alkenes react with *N*-bromosuccinimide (abbreviated NBS) in the presence of light to give products resulting from substitution of hydrogen by bromine at the position *next* to the double bond (the **allylic** position). Cyclohexene, for example, gives 3-bromocyclohexene in 85% yield:

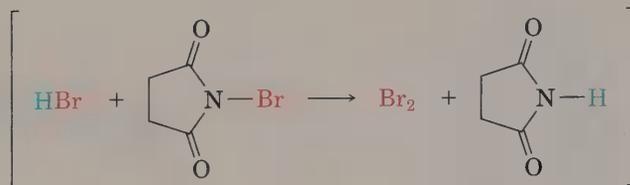
<sup>1</sup>Alfred Wohl (1883–1933); b. Graudentz; Ph.D., Berlin (Hofmann); professor, University of Danzig.



This allylic bromination with NBS looks very similar to the alkane halogenation reaction: In both cases, a C–H bond on a saturated carbon is broken and the hydrogen atom is replaced by halogen. The analogy is a good one, for studies have shown that allylic NBS brominations do in fact occur by a radical pathway. Although the exact mechanism of the reaction is complex, the crucial product-determining step involves abstraction of an allylic hydrogen atom and formation of the corresponding radical. The radical then reacts with  $\text{Br}_2$  to yield the product and a bromine radical, which cycles back into the reaction to carry on the chain. The  $\text{Br}_2$  necessary for reaction with the allyl radical is produced by a concurrent reaction of HBr with NBS:

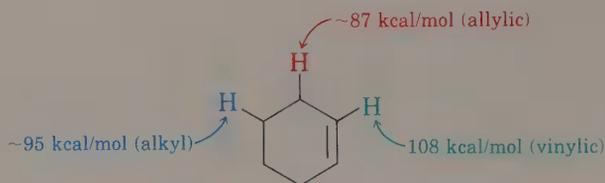


**Allylic radical**



Why does bromination occur exclusively at the position next to the double bond, rather than at one of the other positions? The answer, once again, is found by looking at bond dissociation energies to see the relative stabilities of various kinds of radicals.

There are three kinds of C–H bonds in cyclohexene, and Table 5.4 gives us an idea of their relative strengths. Although a typical secondary alkyl C–H bond has a strength of about 95 kcal/mol (397 kJ/mol), and a typical vinylic C–H bond has a strength of 108 kcal/mol (452 kJ/mol), an *allylic* C–H bond has a strength of only 87 kcal/mol (364 kJ/mol). An allylic radical is therefore more stable than a typical alkyl radical by about 7 kcal/mol (33 kJ/mol).



We can thus expand our stability ordering to include allylic and vinylic radicals:



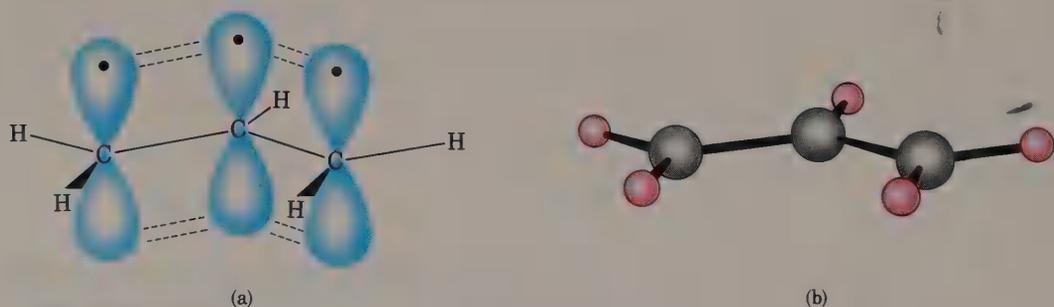
More stable



Less stable

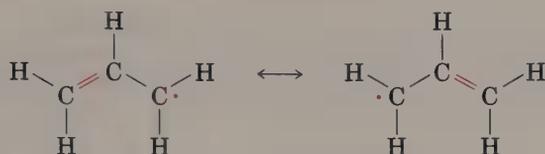
## 10.6 Stability of the Allyl Radical: Resonance

Why are allylic radicals so stable? To see the reason, look at the orbital picture of the allyl radical in Figure 10.3. The radical carbon atom next to the double bond can adopt  $sp^2$  hybridization, placing the unpaired electron in a  $p$  orbital and giving a structure that is electronically symmetrical. The  $p$  orbital on the central carbon can overlap equally well with a  $p$  orbital on either of the two neighboring carbons.



**Figure 10.3** (a) An orbital view of the allyl radical and (b) a computer-generated model. The structure is electronically symmetrical.

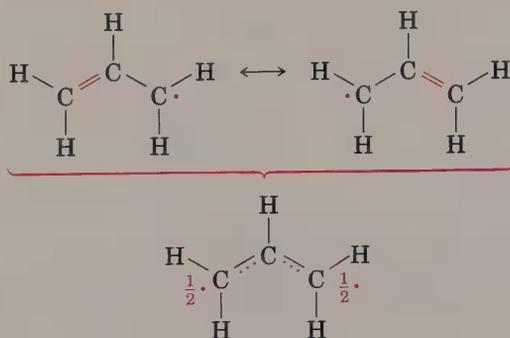
Since the allyl radical is electronically symmetrical, there are two ways in which we can draw it: with the unpaired electron on the left and the double bond on the right, or with the unpaired electron on the right and the double bond on the left. *Neither structure is correct by itself; the true structure of the allyl radical is somewhere in between.*



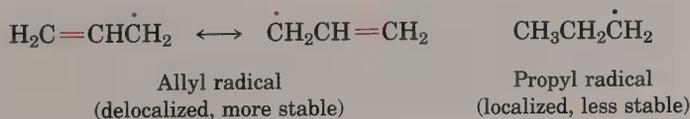
The two structures are called **resonance forms**, and their relationship is indicated by the double-headed arrow between them. The only difference between any pair of resonance forms is the placement of the bonding electrons. *The atoms themselves don't move, but occupy exactly the same places in both resonance forms.*

The best way to think about resonance is to realize that a species like the allyl radical is no different from any other organic substance. An allyl radical doesn't jump back and forth between two resonance forms, spending part of its time looking like one and the rest of its time looking like the other; rather, an allyl radical has a single unchanging structure that we call a **resonance hybrid**.

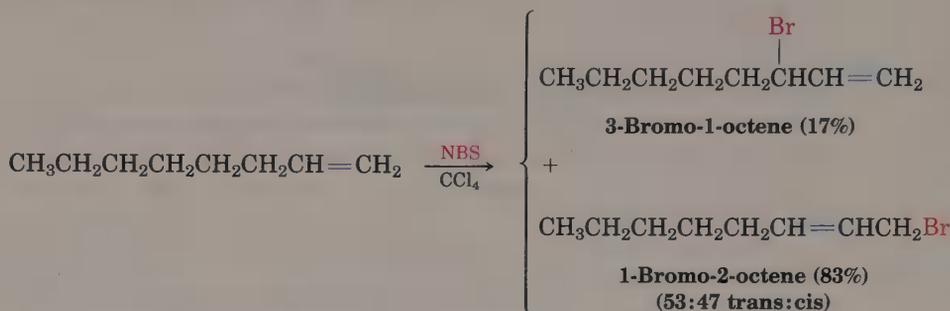
The real difficulty in visualizing the resonance concept is that we can't draw an accurate single picture of a resonance hybrid using familiar Kekulé structures, because the line-bond structures that serve so well to represent most organic molecules don't work well for the allyl radical. The difficulty, however, lies with the *representation* of the structure rather than with the structure itself. We might try to represent the allyl radical by using a dotted line to indicate that the two C–C bonds are equivalent and that each is approximately  $1\frac{1}{2}$  bonds, but such a drawing doesn't help much and will be used infrequently in this book.



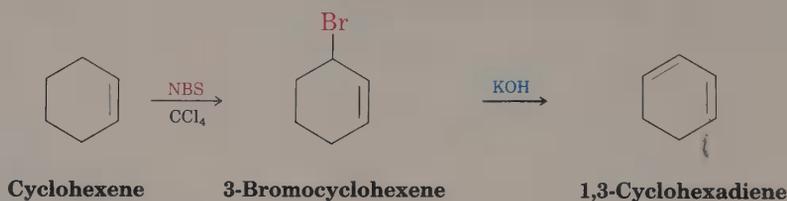
One of the most important postulates of resonance theory is that *the greater the number of possible resonance forms, the greater the stability of the compound*. For example, an allyl radical is a resonance hybrid of two Kekulé structures and is therefore more stable than a typical alkyl radical. In molecular orbital terms, the stability of the allyl radical is due to the fact that the unpaired electron is **delocalized**, or spread out, over an extended  $p_i$  orbital network rather than concentrated at only one site.



In addition to affecting stability, the delocalization of the unpaired electron in the allyl radical has chemical consequences. Since the unpaired electron is spread out over both ends of the pi orbital system, bromination of the allyl radical can occur at either end. Thus, allylic bromination of an unsymmetrical alkene often leads to a mixture of products. For example, bromination of 1-octene gives both 3-bromo-1-octene and 1-bromo-2-octene. Because the intermediate allylic radical in this case is not symmetrical, though, reaction at the two ends is not equally likely and the two products are not formed in equal amounts.

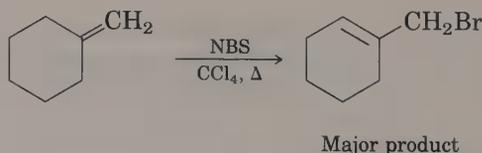


Since product mixtures are formed by bromination of unsymmetrical alkenes, allylic bromination is best carried out on symmetrical alkenes such as cyclohexene. The products of these reactions are particularly useful for conversion into dienes by dehydrohalogenation with base:



PROBLEM .....

- 10.6 The major product of the reaction of methylenecyclohexane with *N*-bromosuccinimide is 1-(bromomethyl)cyclohexene. How do you explain this?



PROBLEM .....

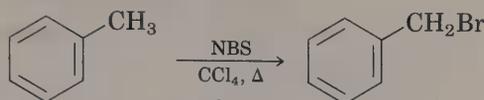
- 10.7 What products would you obtain from reaction of these alkenes with NBS? If more than one product is formed, show the structures of all.

(a) 5-Methylcycloheptene

(b)  $\begin{array}{c} \text{CH}_3 \\ | \\ \text{CH}_3\text{CHCH}=\text{CHCH}_2\text{CH}_3 \end{array}$

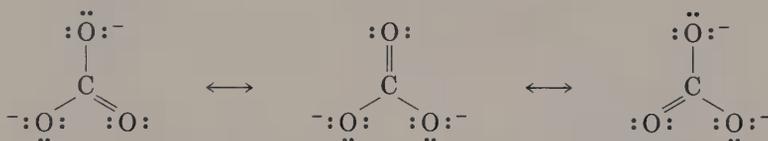
## PROBLEM

- 10.8 Alkylbenzenes such as toluene (methylbenzene) undergo reaction with NBS to give products in which bromine substitution has occurred at the position next to the aromatic ring. How can you account for this result? (See Table 5.4.)

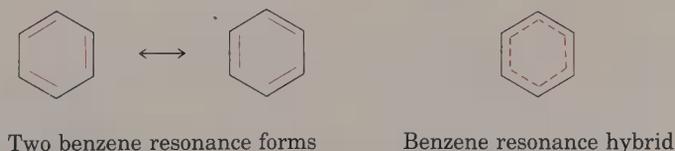


## 10.7 Drawing and Interpreting Resonance Forms

Resonance is an extremely useful concept in chemistry, one that can be used to explain a variety of different phenomena. In inorganic chemistry, for example, the carbonate ion ( $\text{CO}_3^{2-}$ ) is known to have identical bond lengths for its three C–O bonds. Although there is no *single* Lewis or line-bond structure that can account for this equality of C–O bonds, resonance theory accounts for it nicely. The carbonate ion is simply a resonance hybrid of three individual resonance forms. Each of the three oxygens shares the pi electrons and the negative charges equally:



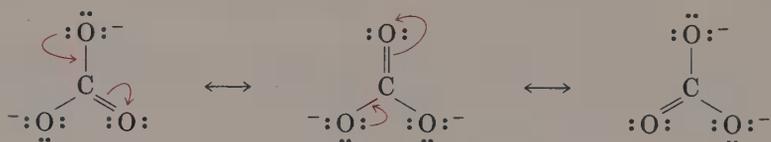
As an example from organic chemistry, we'll see in Chapter 15 that the six C–C bonds in aromatic compounds such as benzene are equivalent, because benzene is a resonance hybrid of two individual resonance forms. Though each individual form seems to imply that benzene has alternating single and double bonds, neither form is correct by itself. The true benzene structure is a hybrid of the two individual forms, and all six carbon–carbon bonds are equivalent:



When first dealing with resonance theory, it's useful to have a set of guidelines that describe how to draw and interpret resonance forms. The following five rules should prove helpful:

**Rule 1** *Resonance forms are imaginary, not real.* The real structure is a composite, or resonance hybrid, of the different forms. Substances such as the allylic radical, the carbonate ion, and benzene are no different from any other substance. They have single, unchanging structures and do not switch back and forth between resonance forms. The only difference with these molecules is in the way they must be represented on paper. Although normal line-bond structures are adequate for representing *most* substances, they're inadequate for representing the bonding in some other substances.

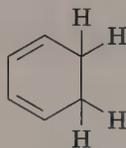
**Rule 2** *Resonance forms differ only in the placement of pi or nonbonding electrons.* Neither the position nor the hybridization of atoms changes from one resonance form to the next. In the carbonate ion, for example, the central carbon atom is  $sp^2$  hybridized in all resonance forms, and the three oxygen atoms remain in exactly the same place in all resonance forms. Only the positions of the pi or nonbonding electrons differ from one form to another:



Similarly, in benzene the pi electrons in the double bonds move, but the carbon atoms remain in place:

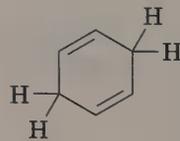


By contrast, two structures such as 1,3-cyclohexadiene and 1,4-cyclohexadiene are *not* resonance structures because they don't have their hydrogen atoms in the same position. Instead, the two dienes are constitutional isomers.



1,3-Cyclohexadiene

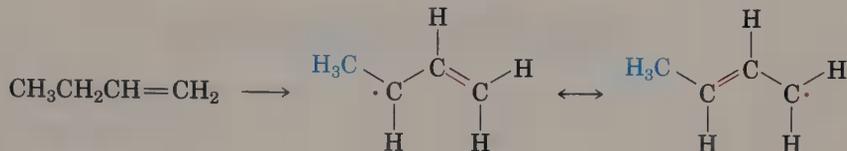
Constitutional isomers,  
*not*  
resonance forms



1,4-Cyclohexadiene

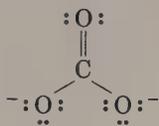
**Rule 3** *Different resonance forms of a substance do not have to be equivalent.* For example, bromination of 1-butene with NBS involves formation of an *unsymmetrical* allylic radical. One end of the delocalized pi electron system has a methyl substituent, and

the other end is unsubstituted. Even though the two resonance forms aren't equivalent, they both still contribute to the overall resonance hybrid.

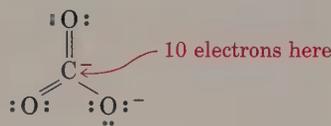


When two resonance forms are nonequivalent, the actual structure of the resonance hybrid is closer to the more stable form than to the less stable form. Thus, we might expect the butenyl radical to look a bit more like a secondary radical than like a primary radical.

**Rule 4** All resonance forms must obey normal rules of valency. Resonance forms are like any other structure: The octet rule still holds. For example, one of the following structures for the carbonate ion is not a valid resonance form because the carbon atom has five bonds and ten electrons:



Carbonate ion

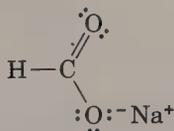


Not a resonance form

**Rule 5** The resonance hybrid is more stable than any single resonance form. In other words, resonance leads to stability. The greater the number of resonance forms possible, the more stable the substance. We've already seen, for example, that an allylic radical is more stable than a normal radical. In a similar manner, we'll see in Chapter 15 that a benzene ring is more stable than a cyclic alkene.

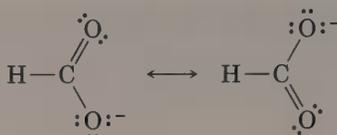
#### PRACTICE PROBLEM.....

Use resonance structures to explain why both C–O bonds of sodium formate have the same length (1.26 Å).



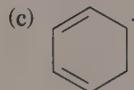
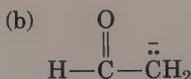
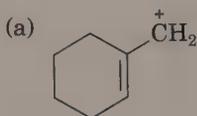
Sodium formate

**Solution** The formate anion is a resonance hybrid of two equivalent resonance forms. The two resonance forms can be drawn by showing the double bond either to the top oxygen or to the bottom oxygen. Only the positions of the electrons are different in the two structures.



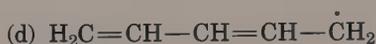
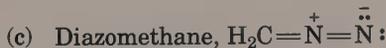
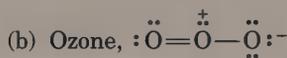
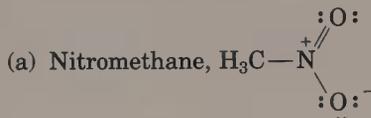
PROBLEM .....

10.9 Draw as many resonance structures as you can for these species:



PROBLEM .....

10.10 Draw as many resonance structures as you can for these species:



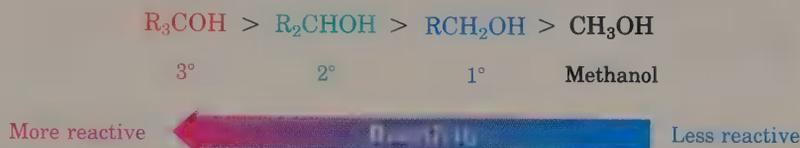
## 10.8 Preparing Alkyl Halides from Alcohols

The most valuable method for the preparation of alkyl halides is their synthesis from alcohols. A great many alcohols are commercially available, and we'll see later that a great many more can be obtained from carbonyl compounds. Because of the importance and generality of the reaction, a variety of different reagents have been used for transforming alcohols into alkyl halides.

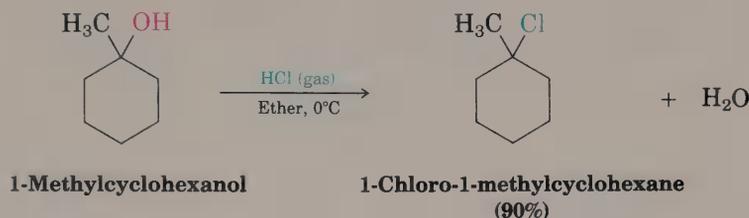
The simplest (but also least generally useful) method for carrying out the conversion of an alcohol to an alkyl halide involves treating the alcohol with HCl, HBr, or HI:



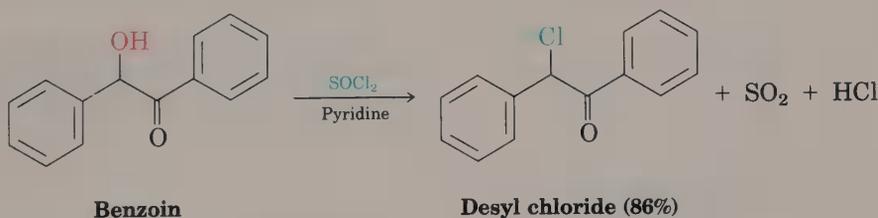
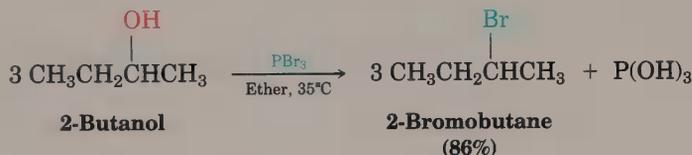
For reasons to be discussed in Section 11.14, the reaction works best when applied to tertiary alcohols,  $\text{R}_3\text{COH}$ . Primary and secondary alcohols also react, but at considerably slower rates and at considerably higher reaction temperatures. Although this is not a problem in simple cases, more complicated molecules are sometimes acid-sensitive and are destroyed by the reaction conditions.



The reaction of HX with a tertiary alcohol is so rapid that it's often carried out simply by bubbling the pure HX gas into a cold ether solution of the alcohol. Reaction is usually complete within a few minutes.



Primary and secondary alcohols are best converted into alkyl halides by treatment with such reagents as thionyl chloride ( $\text{SOCl}_2$ ) or phosphorus tribromide ( $\text{PBr}_3$ ). These reactions, which normally take place readily under mild conditions, are less acidic and less likely to cause acid-catalyzed rearrangements than the HX method is.



As the preceding examples indicate, the yields of these  $\text{PBr}_3$  and  $\text{SOCl}_2$  reactions are generally high. Other functional groups such as ethers, carbonyls, and aromatic rings don't usually interfere. We'll look at how these substitution reactions take place in the next chapter.

PROBLEM.....

10.11 How would you prepare these alkyl halides from the appropriate alcohols?

(a) 2-Chloro-2-methylpropane

(b) 2-Bromo-4-methylpentane

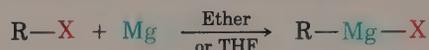
(c)  $\text{BrCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)_2$

(d)  $\text{CH}_3\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CCl}(\text{CH}_3)_2$

.....

## 10.9 Reactions of Alkyl Halides: Grignard Reagents

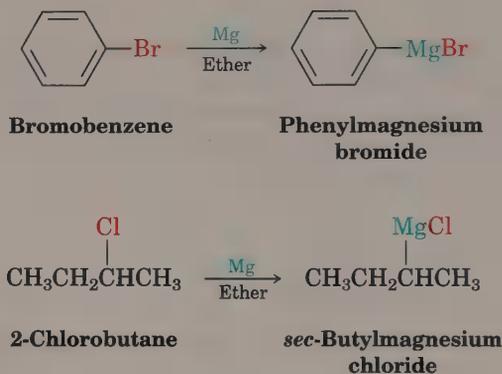
Organic halides of widely varying structure—alkyl, aryl, and vinylic—react with magnesium metal in ether or tetrahydrofuran (THF) solvent to yield organomagnesium halides,  $\text{RMgX}$ . These products, named **Grignard reagents** after their discoverer, Victor Grignard,<sup>2</sup> are examples of organo-metallic compounds since they contain a carbon–metal bond.



where R = 1°, 2°, or 3° alkyl, aryl, or alkenyl

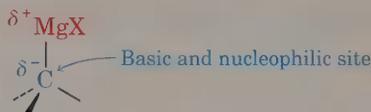
X = Cl, Br, or I

For example:



Many different kinds of organohalides form Grignard reagents. Steric hindrance in the halide doesn't appear to be a factor in the formation of Grignard reagents, since 1°, 2°, and 3° alkyl halides all react with ease. Aryl and alkenyl halides also react, although it's best to use THF as solvent for these cases. The halogen may be Cl, Br, or I, although chlorides are somewhat less reactive than bromides and iodides. Organofluorides rarely react with magnesium.

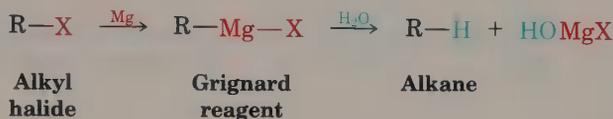
As you might expect from the discussion of electronegativity and bond polarity in Section 5.5, the carbon–magnesium bond is highly polarized, making the carbon atom both nucleophilic and basic:



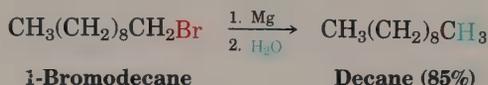
<sup>2</sup>François Auguste Victor Grignard (1871–1935); b. Cherbourg, France; professor, University of Nancy, Lyons; Nobel Prize (1912).

In a formal sense, a Grignard reagent can be considered to be a carbon anion or **carbanion**—the magnesium salt of a hydrocarbon acid. It's more accurate, though, to view Grignard reagents as containing a highly polar covalent C–Mg bond, rather than an ionic bond between C<sup>−</sup> and <sup>+</sup>MgX.

Because of their nucleophilic/basic character, Grignard reagents react both with acids and with a wide variety of electrophiles. For example, Grignard reagents react with proton donors (Brønsted acids) such as H<sub>2</sub>O, ROH, RCOOH, or RNH<sub>2</sub> to yield hydrocarbons. The overall sequence is a useful synthetic method for converting an alkyl halide into an alkane (R–X → R–H):



For example,



We'll see many more uses of Grignard reagents in later chapters.

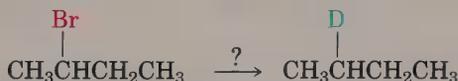
PROBLEM.....

10.12 How strong a base would you expect a Grignard reagent to be? Look at Tables 8.1 and 8.2, and then predict whether the following reactions are likely to occur.



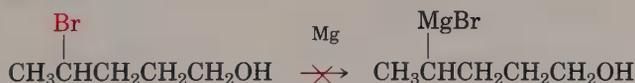
PROBLEM.....

10.13 An important advantage of alkyl halide reduction via Grignard reagents is that the sequence can be used to introduce deuterium into a specific site in a molecule. How might you do this?



PROBLEM.....

10.14 Why do you suppose it's not possible to prepare a Grignard reagent from a bromo alcohol such as 4-bromo-1-pentanol?

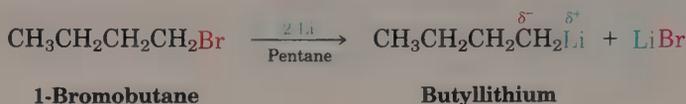


Give another example of a molecule that is unlikely to form a Grignard reagent.

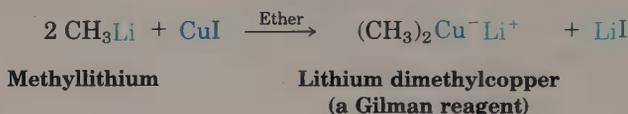
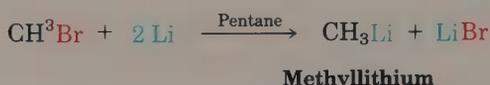
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## 10.10 Organometallic Coupling Reactions

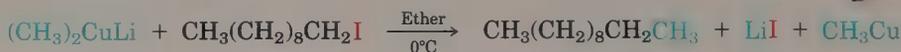
Several other organometallic reagents can be prepared in a manner similar to that of Grignard reagents. For example, alkyllithium reagents can be prepared by the reaction of an alkyl halide with lithium metal:



Alkyllithiums are both nucleophiles and bases, and their chemistry is similar in many respects to that of the alkylmagnesium halides. One of the most valuable reactions of alkyllithiums is their use in preparing lithium diorganocopper reagents, or **Gilman<sup>3</sup> reagents**.

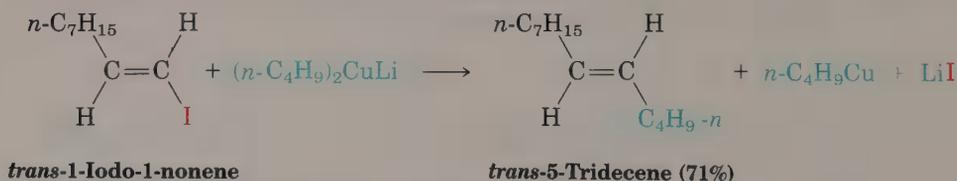


Gilman reagents are easily prepared by reaction of an alkyllithium with cuprous iodide, CuI, in ether solvent. Though rather unstable, they have the remarkable ability of undergoing organometallic *coupling* reactions with alkyl bromides and iodides (but not fluorides). One of the alkyl groups from the Gilman reagent replaces the halogen from alkyl halide, resulting in the formation of a hydrocarbon with a new carbon-carbon bond. Lithium dimethylcopper, for example, reacts with 1-iododecane to give undecane in 90% yield:

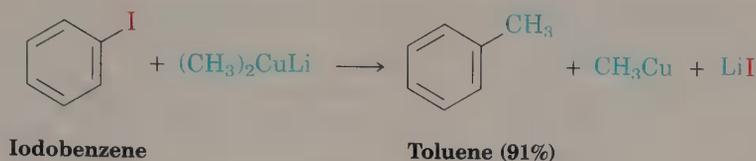


Lithium dimethylcopper    1-Iododecane Undecane (90%)

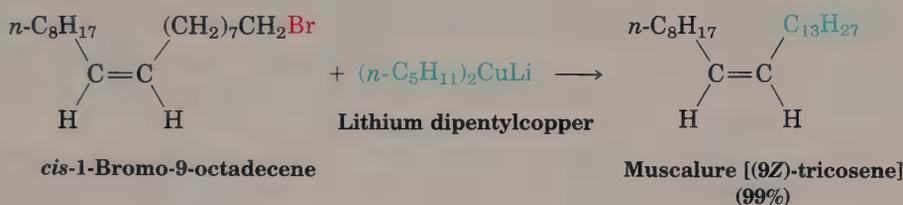
This organometallic coupling reaction is extremely versatile and very useful in organic synthesis. As the following examples indicate, the coupling reaction can be carried out on aryl and vinylic halides as well as on alkyl halides.



<sup>3</sup>Henry Gilman (1893–1986); b. Boston; Ph.D. (1918) Harvard (Kohler); professor, Iowa State University (1923–1986).



An organocopper coupling reaction is even carried out on a commercial scale to synthesize *muscalure*, (*9Z*)-tricosene, the sex attractant secreted by the common housefly. Minute amounts of this insect hormone, or **pheromone**, greatly increase the lure of insecticide-treated fly bait and provide a species-specific means of insect control. Coupling of *cis*-1-bromo-9-octadecene with lithium dipentylcopper is used industrially to produce *muscalure* in 100 lb batches.



Although the details of the mechanism by which coupling occurs are not fully understood, radicals may be involved. The coupling is not a typical polar nucleophilic substitution reaction of the sort considered in the next chapter.

**PROBLEM** .....

- 10.15** How would you prepare these compounds using an organocopper coupling reaction at some point in the scheme? More than one step is required in each case.
- 3-Methylcyclohexene from cyclohexene
  - Octane from 1-bromobutane
  - Decane from 1-pentene
- .....

## 10.11 Summary and Key Words

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**Alkyl halides** are compounds containing halogen bonded to a saturated,  $sp^3$ -hybridized carbon atom. The C–X bond is polar, and alkyl halides can therefore behave as electrophiles.

Alkyl halides can be prepared by **radical halogenation** of alkanes, but this method is of little general value since mixtures of products usually result. The reactivity order of alkanes toward halogenation is identical to the stability order of radicals: tertiary > secondary > primary. According to the Hammond postulate, the more stable radical intermediate is formed faster because the transition state leading to it is more stable.

Alkyl halides can also be prepared from alkenes. Alkenes react with *N*-bromosuccinimide (NBS) to give the product of **allylic bromination**. The

NBS bromination of alkenes is a complex process that takes place through an intermediate allyl radical. Allyl radicals are stabilized by **resonance** and can be drawn in two different ways, neither of which is correct by itself. The true structure of the allyl radical is best described as a composite, or **resonance hybrid**, of the two individual resonance forms. The only difference between the two resonance structures is in the location of bonding electrons: The nuclei remain in the same places in both structures.

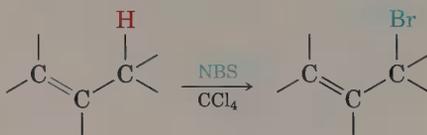


Alcohols react with HX to form alkyl halides, but the reaction works well only for tertiary alcohols,  $\text{R}_3\text{COH}$ . Primary and secondary alkyl halides are normally prepared from alcohols using either  $\text{SOCl}_2$  or  $\text{PBr}_3$ . Alkyl halides react with magnesium in ether solution to form organomagnesium halides, or **Grignard reagents**,  $\text{RMgX}$ . Since Grignard reagents are both nucleophilic and basic, they react with Brønsted acids to form hydrocarbons. The overall result of Grignard formation and protonation is the conversion of an alkyl halide into an alkane ( $\text{RX} \rightarrow \text{RMgX} \rightarrow \text{RH}$ ).

Alkyl halides also react with lithium metal to form **organolithium reagents**,  $\text{RLi}$ . In the presence of  $\text{CuI}$ , these form diorganocoppers, or **Gilman reagents**,  $\text{R}_2\text{CuLi}$ . Gilman reagents react with alkyl halides to yield coupled hydrocarbon products.

## 10.12 Summary of Reactions

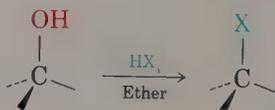
1. Preparation of alkyl halides
  - a. From alkenes by allylic bromination (Section 10.5)



- b. From alkenes by addition of HBr and HCl (Sections 6.9 and 6.10)

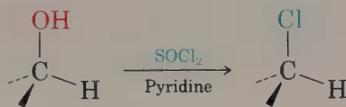


- c. From alcohols
    - (1) Reaction with HX, where X = Cl, Br, or I (Section 10.8)

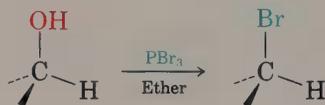


Reactivity order:  $3^\circ > 2^\circ > 1^\circ$

- (2) Reaction of 1° and 2° alcohols with  $\text{SOCl}_2$  in pyridine (Section 10.8)

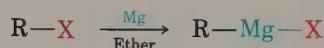


- (3) Reaction of 1° and 2° alcohols with  $\text{PBr}_3$  (Section 10.8)



## 2. Reactions of alkyl halides

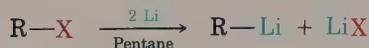
- a. Grignard reagent formation (Section 10.9)



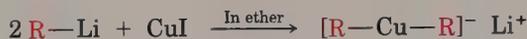
where X = Br, Cl, or I

R = 1°, 2°, 3° alkyl, aryl, or vinylic

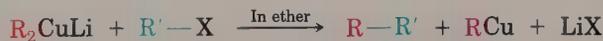
- b. Diorganocopper (Gilman reagent) formation (Section 10.10)



where R = 1°, 2°, 3° alkyl, aryl, or vinylic



- c. Organometallic coupling (Section 10.10)



- d. Conversion of alkyl halides to alkanes (Section 10.9)



## ADDITIONAL PROBLEMS .....

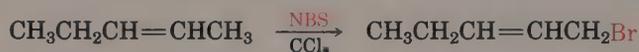
**10.16** Name these alkyl halides according to IUPAC rules:

- (a)  $(\text{CH}_3)_2\text{CHCHBrCHBrCH}_2\text{CH}(\text{CH}_3)_2$       (b)  $\text{CH}_3\text{CH}=\text{CHCH}_2\text{CHICH}_3$   
 (c)  $(\text{CH}_3)_2\text{CBrCH}_2\text{CHClCH}(\text{CH}_3)_2$       (d)  $\text{CH}_3\text{CH}_2\text{CH}(\text{CH}_2\text{Br})\text{CH}_2\text{CH}_2\text{CH}_3$   
 (e)  $\text{ClCH}_2\text{CH}_2\text{CH}_2\text{C}\equiv\text{CCH}_2\text{Br}$

10.17 Draw structures corresponding to these IUPAC names:

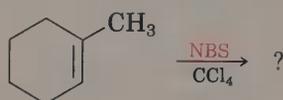
- (a) 2,3-Dichloro-4-methylhexane                      (b) 4-Bromo-4-ethyl-2-methylhexane  
 (c) 3-Iodo-2,2,4,4-tetramethylpentane              (d) *cis*-1-Bromo-2-ethylcyclopentane

10.18 A chemist requires a large amount of 1-bromo-2-pentene as starting material for a synthesis. She finds a supply of 2-pentene in the stockroom and decides to carry out an NBS allylic bromination reaction:



What is wrong with this synthesis plan? What side products would form in addition to the desired product?

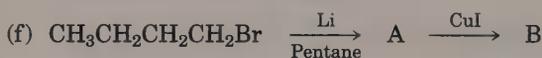
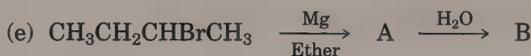
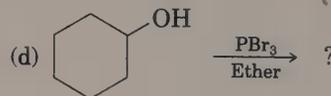
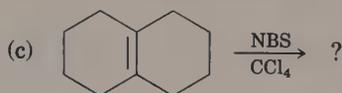
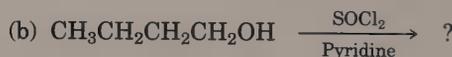
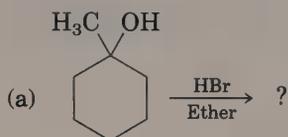
10.19 What product(s) would you expect from the reaction of 1-methylcyclohexene with NBS? Would you use this reaction as part of a synthesis?



10.20 How would you prepare the following compounds, starting with cyclopentene and any other reagents needed?

- (a) Chlorocyclopentane                                      (b) Methylcyclopentane  
 (c) 3-Bromocyclopentene                                  (d) Cyclopentanol  
 (e) Cyclopentylcyclopentane                              (f) 1,3-Cyclopentadiene

10.21 Predict the product(s) of these reactions:

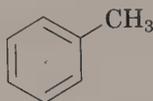


10.22 Draw and name the monochlorination products you might obtain by radical chlorination of 2-methylpentane. Which of the products are chiral? Are any of the products optically active?

10.23 (*S*)-3-Methylhexane undergoes radical bromination to yield 3-bromo-3-methylhexane as the major product. Is the product chiral? Is it optically active? What stereoisomers are produced and in what ratio?

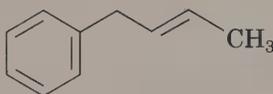
10.24 Assume that you have carried out a radical chlorination reaction on (*R*)-2-chloropentane and have isolated (in low yield) 2,4-dichloropentane. How many stereoisomers of the product are formed and in what ratio? Are any of the isomers optically active?

- 10.25 Calculate  $\Delta H^\circ$  for the reactions of  $\text{Cl}\cdot$  and  $\text{Br}\cdot$  with  $\text{CH}_4$ , and then draw a reaction energy diagram showing both processes. Which reaction is likely to be faster?
- 10.26 Table 5.4 shows that the methyl C–H bond of toluene is 13 kcal/mol (54 kJ/mol) weaker than the C–H bond of ethane. In other words, the  $\text{C}_6\text{H}_5\text{CH}_2\cdot$  (benzyl) radical is 13 kcal/mol more stable than the  $\cdot\text{CH}_2\text{CH}_3$  radical. Draw as many resonance structures as you can for the benzyl radical.



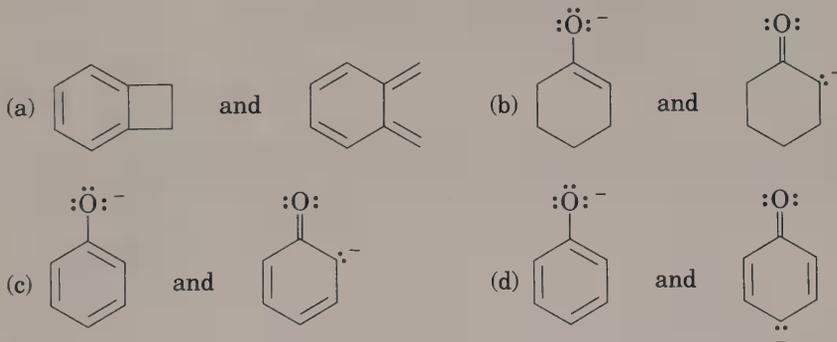
Toluene

- 10.27 What product(s) would you expect from the reaction of 1-phenyl-2-butene with NBS? Explain.

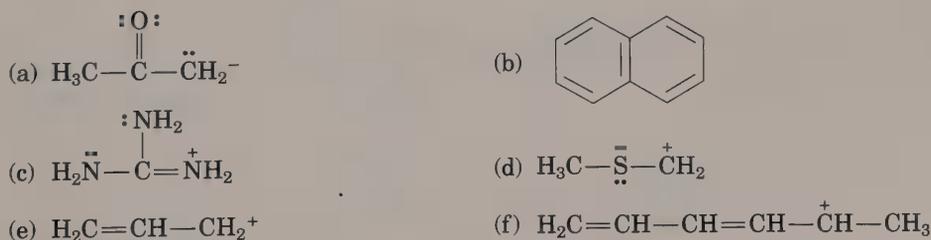


1-Phenyl-2-butene

- 10.28 Which of the following pairs of structures represent resonance forms?



- 10.29 Draw as many resonance structures as you can for these species:



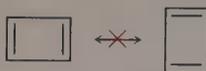
- 10.30 Which of the following two resonance structures would you expect to contribute most to the resonance hybrid? Explain.



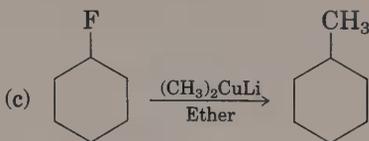
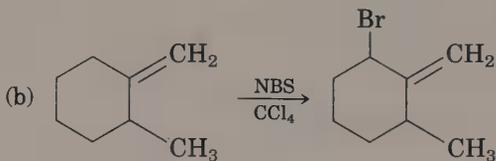
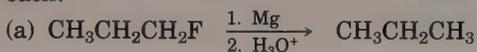
10.31 How would you carry out these syntheses?

- Butylcyclohexane from cyclohexene
- Butylcyclohexane from cyclohexanol
- Butylcyclohexane from cyclohexane

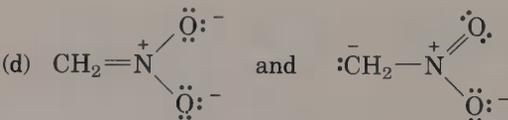
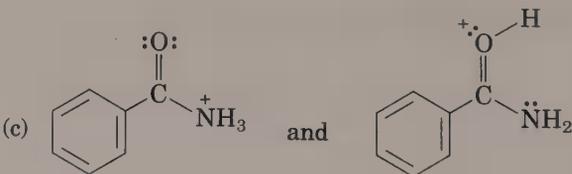
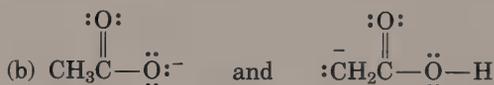
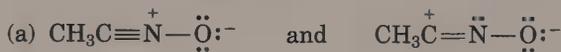
10.32 Cyclobutadiene is a rectangular molecule with two shorter single bonds and two longer double bonds. Why do the following structures *not* represent resonance forms?



10.33 The syntheses shown here are unlikely to occur as written. What is wrong with each?



10.34 Which of the following pairs represent resonance structures?



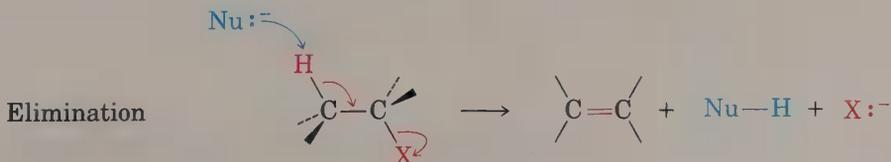
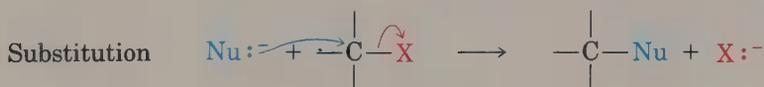
# Chapter 11

## Reactions of Alkyl Halides: Nucleophilic Substitutions and Eliminations

We saw in the preceding chapter that the carbon-halogen bond in alkyl halides is polar and the carbon atom is electron-poor. Thus, alkyl halides are electrophiles, and much of their chemistry involves polar reactions with electron-rich nucleophiles and bases.



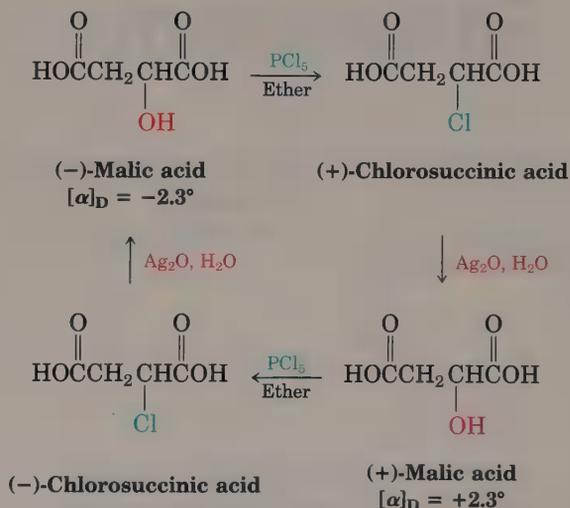
Alkyl halides do one of two things when they react with a nucleophile: Either they undergo *substitution* of the X group by the nucleophile or they undergo *elimination* of HX to yield an alkene.



Nucleophilic substitution and base-induced elimination are two of the most important reactions in organic chemistry. It's now time to take a close look at them to see how these reactions occur, what factors are involved, and how we can control the reactions for purposes of organic synthesis.

## 11.1 The Discovery of the Walden Inversion

In 1896 the German chemist Paul Walden<sup>1</sup> reported a remarkable discovery. He found that the pure enantiomeric (+)- and (-)-malic acids could be interconverted by a series of simple substitution reactions. When Walden treated (-)-malic acid with  $\text{PCl}_5$ , he isolated dextrorotatory (+)-chlorosuccinic acid. This, on treatment with wet silver oxide, gave (+)-malic acid. Similarly, reaction of (+)-malic acid with  $\text{PCl}_5$  gave levorotatory (-)-chlorosuccinic acid, which was converted into (-)-malic acid when treated with wet silver oxide. The full cycle of reactions reported by Walden is shown in Figure 11.1.



**Figure 11.1** Walden's cycle of reactions interconverting (+)- and (-)-malic acids.

At the time, the results were astonishing. The eminent chemist Emil Fischer called Walden's discovery "the most remarkable observation made in the field of optical activity since the fundamental observations of Pasteur."

<sup>1</sup>Paul Walden (1863–1957); b. Latvia; Ph.D., Leipzig; student and professor, Riga Polytechnic, Russia (1882–1919); professor, University of Rostock, University of Tübingen, Germany.

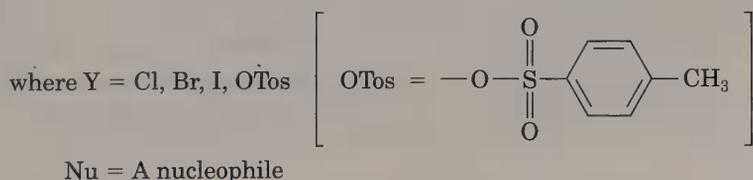
Since (–)-malic acid was being converted into (+)-malic acid, *some reactions in the cycle must have occurred with a change in configuration at the stereogenic center*. But which ones, and how? (Recall that the direction of light rotation and the absolute configuration of a molecule aren't directly related. You can't tell by looking at the sign of rotation whether or not a change in configuration has occurred during a reaction.)

Today, we refer to the transformations taking place in Walden's cycle as **nucleophilic substitution reactions** because each step involves the substitution of one nucleophile (chloride ion,  $\text{Cl}^-$ , or hydroxide ion,  $\text{HO}^-$ ) by another. Nucleophilic substitution reactions are one of the most important general reaction types in organic chemistry.

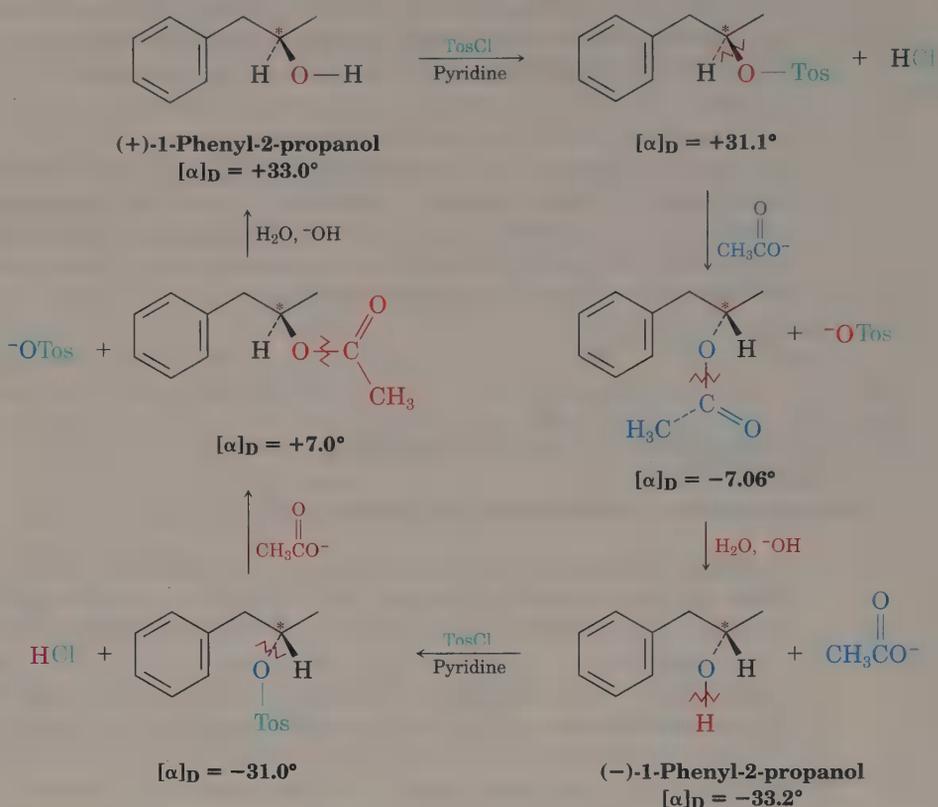
## 11.2 Stereochemistry of Nucleophilic Substitution

Although Walden realized that changes in configuration must have taken place during his reaction cycle, he didn't know at which steps the changes occurred. In the 1920s, Joseph Kenyon<sup>2</sup> and Henry Phillips began a series of investigations to elucidate the mechanism of nucleophilic substitution reactions and to find out how inversions of configuration occur. They recognized that the presence of the carboxylic acid group in Walden's work on malic acid may have led to complications, and they therefore carried out their own work on simpler cases. (In fact, the particular sequence of reactions studied by Walden *is* unusually complex for reasons we won't go into. The crucial point the sequence raised, however, is the idea that changes in three-dimensional configuration can evidently occur in organic reactions.)

Among the reactions studied by Kenyon and Phillips was one that interconverted the two enantiomers of 1-phenyl-2-propanol (Figure 11.2). Although this particular series of reactions involves nucleophilic substitution of an alkyl toluenesulfonate (a *tosylate*) rather than an alkyl halide, exactly the same type of reaction is involved as that studied by Walden. For all practical purposes, the *entire* tosylate group acts as if it were simply a halogen substituent:

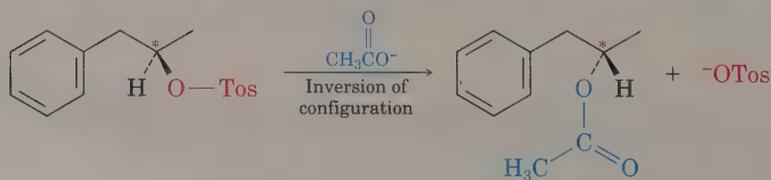


<sup>2</sup>Joseph Kenyon (1885–1961); b. Blackburn, England; D.Sc. (London), 1914; British Dyestuffs Corp. (1916–1920); Battersea Polytechnic, London (1920–1950).



**Figure 11.2** A Walden cycle on 1-phenyl-2-propanol. Stereogenic centers are marked by asterisks, and the bonds broken in each reaction are indicated by wavy lines.

In the three-step reaction sequence shown in Figure 11.2, (+)-1-phenyl-2-propanol is converted into its (–) enantiomer, and at least one of the three steps must therefore involve an *inversion* (change) of configuration at the stereogenic center. The first step, formation of a toluenesulfonate, is known to occur by breaking the O–H bond of the alcohol rather than the C–O bond to the chiral carbon, and the configuration around carbon is therefore unchanged. Similarly, it can be shown (Section 21.7) that the third step, hydroxide ion cleavage of the acetate, also takes place without breaking the C–O bond at the stereogenic center. Thus, inversion can't occur in the third step. *The inversion of stereochemical configuration must therefore take place in the second step, the nucleophilic substitution of tosylate ion by acetate ion.*



From this and nearly a dozen other series of reactions, Kenyon and Phillips concluded that the nucleophilic substitution reaction of primary and secondary alkyl halides and tosylates always proceeds with inversion of configuration.

PROBLEM .....

- 11.1 What product would you expect to obtain from reaction of (*S*)-2-bromohexane with sodium acetate,  $\text{Na}^+ \text{ } ^-\text{OOCCH}_3$ ? Show the stereochemistry of both starting material and product.
- .....

## 11.3 Kinetics of Nucleophilic Substitution

Chemists often speak about a reaction as being fast or slow. The exact rate at which a starting material reacts to give product is called the **reaction rate** and can often be measured. The determination of reaction rates and of how those rates depend on reagent concentrations is a powerful tool for probing reaction mechanisms. Let's see what can be learned about the nucleophilic substitution reaction by a study of reaction rates.

In all chemical reactions, there is a direct relationship between reaction rate and reagent concentrations. When we measure this relationship, we measure the **kinetics** of the reaction. For example, let's look at some factors that influence the rate of a simple nucleophilic substitution: the reaction of bromomethane with hydroxide ion to yield methanol.



At a given concentration of reagents, the reaction occurs at a certain rate. If we double the concentration of hydroxide ion, the frequency of encounter between the reagents is also doubled, and we might therefore predict that the reaction rate will double. Similarly, if we double the concentration of bromomethane, we might expect that the reaction rate will again double. This behavior is exactly what is found. We call such a reaction, in which the rate is linearly dependent on the concentrations of two reagents, a **second-order reaction**. Mathematically, we can express this second-order dependence of the nucleophilic substitution reaction by setting up a **rate equation**:

$$\begin{aligned} \text{Reaction rate} &= \text{Rate of disappearance of starting material} \\ &= k \times [\text{RX}] \times [\text{ } ^-\text{OH}] \end{aligned}$$

where  $[\text{RX}] = \text{CH}_3\text{Br}$  concentration

$[\text{ } ^-\text{OH}] = \text{ } ^-\text{OH}$  concentration

$k = \text{A constant value}$

This equation says that the reaction rate is the same as the rate of disappearance of starting material and is equal to a coefficient,  $k$ , times the alkyl halide concentration times the hydroxide ion concentration. The constant  $k$  is called the **rate coefficient** for the reaction and has units of liters per mole second (L/mol sec). The rate equation simply says that as either  $[RX]$  or  $[^-OH]$  changes, the rate of the reaction changes accordingly. If the alkyl halide concentration is doubled, the reaction rate doubles; if the alkyl halide concentration is halved, the reaction rate is halved.

## 11.4 The $S_N2$ Reaction

At this point, two important pieces of information have been obtained about the nature of nucleophilic substitution reactions on primary and secondary alkyl halides and tosylates:

1. These reactions always occur with inversion of stereochemistry at the stereogenic carbon center.
2. These reactions show second-order kinetics and follow the rate law:

$$\text{Rate} = k \times [RX] \times [\text{Nu}^-]$$

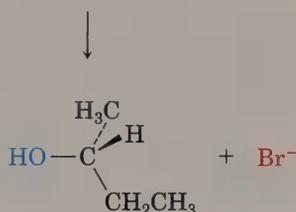
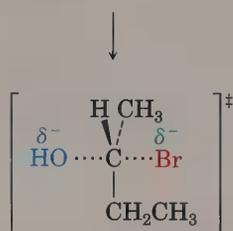
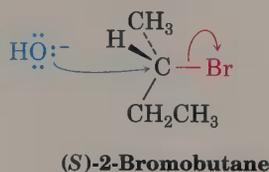
What mechanism will account for the reaction stereochemistry and second-order kinetics of these reactions? An explanation was first advanced in 1937 when Hughes<sup>3</sup> and Ingold formulated the mechanism of what they called the  **$S_N2$  reaction**—shorthand for *substitution, nucleophilic, bimolecular*. (**Bimolecular** means that two molecules, nucleophile and alkyl halide, take part in the step whose kinetics are measured.)

The essential feature of the  $S_N2$  reaction mechanism formulated by Hughes and Ingold is that the reaction takes place in a single step without intermediates when the incoming nucleophile attacks the substrate from a position  $180^\circ$  away from the departing group. As the nucleophile comes in on one side and bonds to the chiral carbon, the halide or tosylate departs from the other side, and the stereochemical configuration of the molecule inverts. The process is shown in Figure 11.3 for the reaction of (*S*)-2-bromobutane with hydroxide ion, leading to (*R*)-2-butanol.

We can picture the reaction as occurring when an electron pair on the nucleophile,  $\text{Nu}^-$ , forces out the leaving group,  $\text{Y}^-$ , with the electron pair from the former  $\text{C}-\text{Y}$  bond. This occurs through a transition state in which the new  $\text{Nu}-\text{C}$  bond is partially forming at the same time that the old  $\text{C}-\text{Y}$  bond is partially breaking, and in which the negative charge is shared by both the incoming nucleophile and the outgoing leaving group. The transition state for this inversion must have the remaining three bonds to carbon in a planar arrangement, as shown in Figure 11.4.

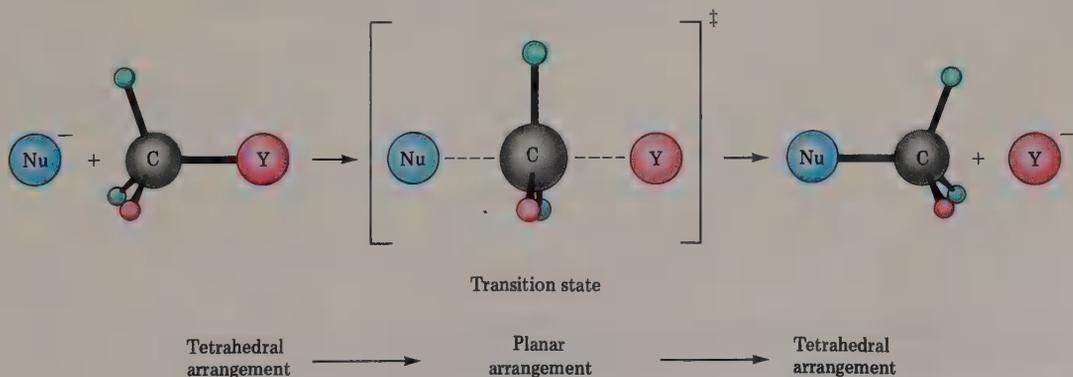
<sup>3</sup>Edward David Hughes (1906–1963); b. Criccieth, North Wales; Ph.D., Wales (Watson); D.Sc., London (Ingold); professor, University College, London (1930–1963).

The nucleophile <sup>-</sup>OH uses its lone-pair electrons to attack the alkyl halide carbon 180° away from the departing halogen. This leads to a transition state with a partially formed C–OH bond and a partially broken C–Br bond.

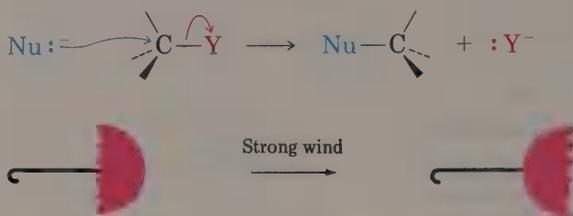


The stereochemistry at carbon is inverted as the C–OH bond forms fully and the bromide ion departs with the electron pair from the former C–Br bond.

**Figure 11.3** The mechanism of the S<sub>N</sub>2 reaction. The reaction takes place in a single step when the incoming nucleophile approaches from a direction 180° away from the departing halide ion. The stereochemistry at carbon is inverted from *S* in the starting material to *R* in the product.



**Figure 11.4** The planar transition state of an S<sub>N</sub>2 reaction.



**Figure 11.5** The inversion of a stereogenic center during an  $S_N2$  reaction is similar to the inversion of an umbrella in a strong wind.

The mechanism proposed by Hughes and Ingold is fully consistent with experimental results, explaining both stereochemical and kinetic data. Thus, the requirement for back-side attack of the entering nucleophile from a direction  $180^\circ$  away from the departing Y group causes the stereochemistry of the substrate to invert, much like an umbrella turning inside out in the wind (Figure 11.5). The Hughes–Ingold mechanism also explains why second-order kinetics are found for  $S_N2$  reactions: The reaction occurs in a single step that involves *both* alkyl halide and nucleophile. Two molecules are involved in the step whose rate is measured.

**PROBLEM** .....

- 11.2 What product would you expect to obtain from reaction of NaOH with (*R*)-2-bromobutane? Formulate the reaction showing the stereochemistry of both starting material and product.

**PROBLEM** .....

- 11.3 A further piece of evidence in support of the requirement for back-side  $S_N2$  displacement is the finding that the following alkyl bromide does not undergo a substitution reaction with hydroxide ion. Can you suggest a reason for the lack of reactivity? Making a molecular model should be helpful.



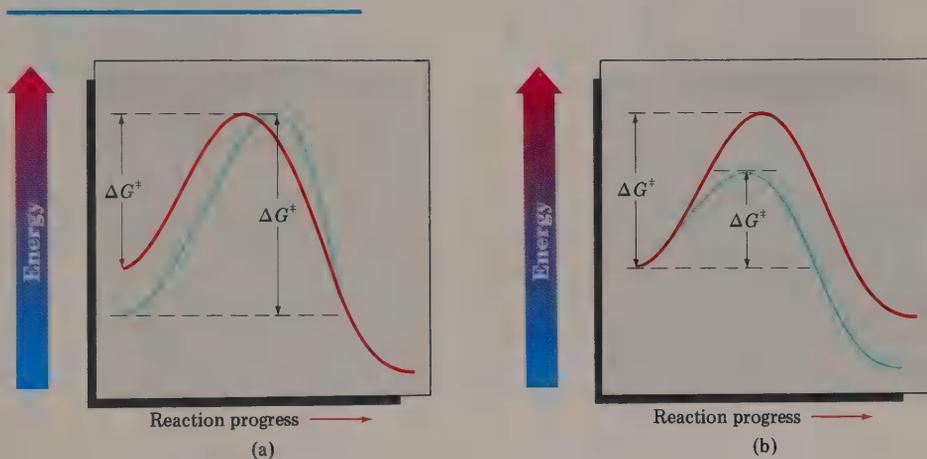
## 11.5 Characteristics of the $S_N2$ Reaction

We now have a good picture of how  $S_N2$  reactions occur, but we also need to see how these substitutions can be used and what variables affect them. Some  $S_N2$  reactions are fast and some are slow; some take place in high

yield, and others in low yield. Understanding the different factors involved can be of tremendous value to chemists. Let's begin by reviewing what we've learned about reaction rates in general.

The rate of a chemical reaction is determined by  $\Delta G^\ddagger$ , the energy difference between reactant (ground state) and transition state. A change in reaction conditions can affect the magnitude of  $\Delta G^\ddagger$  in two ways: either by changing the *reactant* energy level or by changing the *transition-state* energy level. If the reactant energy level is lowered,  $\Delta G^\ddagger$  increases and reaction rate decreases; conversely, if the reactant energy level is raised,  $\Delta G^\ddagger$  decreases and reaction rate increases, as can be seen in Figure 11.6(a).

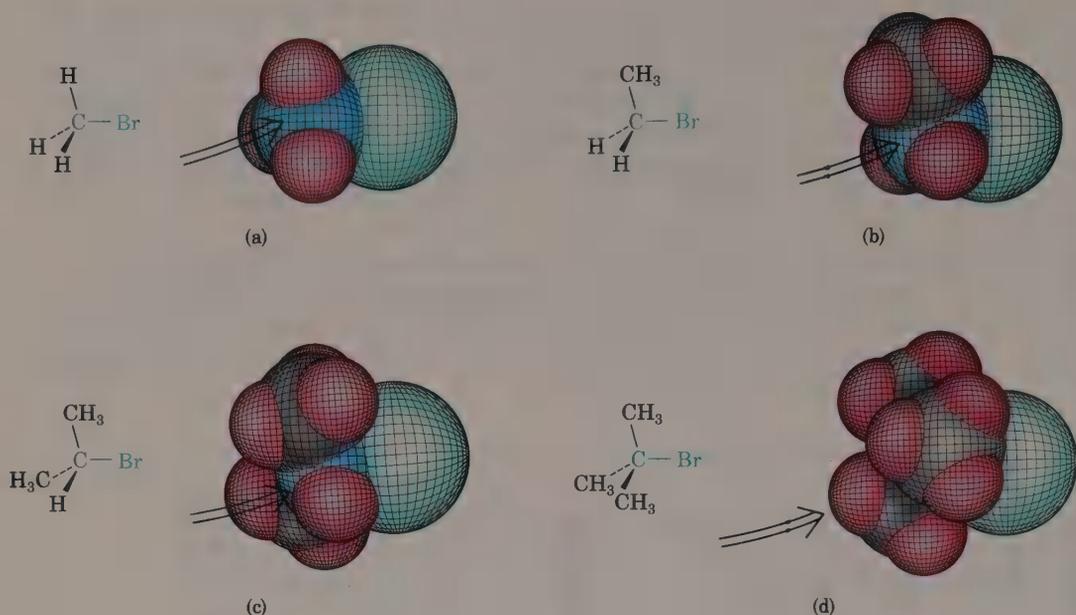
Similarly, stabilization of the transition state lowers  $\Delta G^\ddagger$  and raises reaction rate, whereas destabilization of the transition state raises  $\Delta G^\ddagger$  and lowers reaction rate [Figure 11.6(b)]. We'll see examples of all these effects as we look at  $S_N2$  reaction variables.



**Figure 11.6** The effect of changes in reactant and transition-state energy levels on reaction rate. (a) Higher reactant energy level corresponds to faster reaction (lower  $\Delta G^\ddagger$ ). (b) Higher transition-state energy level corresponds to slower reaction (higher  $\Delta G^\ddagger$ ).

## THE SUBSTRATE: STERIC EFFECTS IN THE $S_N2$ REACTION

The first  $S_N2$  reaction variable to look at is the bulk of the alkyl halide. We've said that the transition state for an  $S_N2$  reaction involves partial bonding between the attacking nucleophile and the substrate. It therefore seems reasonable that a hindered, bulky substrate should prevent easy approach of an incoming nucleophile and should have a more difficult time reaching the transition state. In other words, sterically bulky substrates, in which the carbon atom is "shielded" from attack by the incoming nucleophile, should react more slowly than less hindered substrates because the transition state for their reaction is sterically hindered and higher in energy (Figure 11.7).



**Figure 11.7** Steric hindrance to the  $S_N2$  reaction. As the computer-generated space-filling models indicate, the carbon atom in (a) bromomethane is readily accessible, resulting in a fast  $S_N2$  reaction. The carbon atoms in (b) bromoethane (primary), (c) 2-bromopropane (secondary), and (d) 2-bromo-2-methylpropane (tertiary) are successively less accessible, resulting in successively slower  $S_N2$  reactions.

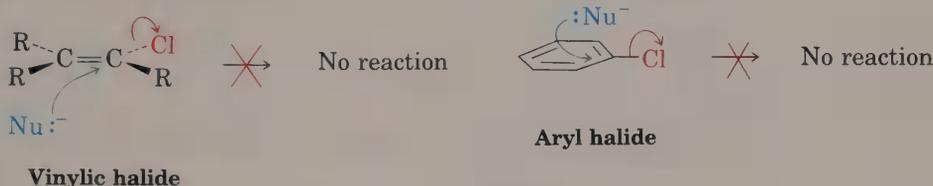
As Figure 11.7 shows, the difficulty of nucleophilic attack increases as the three substituents bonded to the halo-substituted carbon atom increase in size. The relative reactivities for some different substrates are as follows:

	$\begin{array}{c} \text{H} \\   \\ \text{H}-\text{C}-\text{Br} \\   \\ \text{H} \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{CH}_3-\text{C}-\text{Br} \\   \\ \text{H} \end{array}$	$\begin{array}{c} \text{H} \\   \\ \text{CH}_3-\text{C}-\text{Br} \\   \\ \text{CH}_3 \end{array}$	$\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3-\text{C}-\text{CH}_2-\text{Br} \\   \\ \text{CH}_3 \end{array}$	$\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3-\text{C}-\text{Br} \\   \\ \text{CH}_3 \end{array}$
Relative reactivity	(methyl) 3,000,000	(primary) 100,000	(secondary) 2,500	(neopentyl) 1.0	(tertiary) < 1
	<div style="display: flex; align-items: center; justify-content: center;"> <span style="margin-right: 10px;">More reactive</span> <span style="margin-left: 10px;">Less reactive</span> </div>				

Methyl halides are by far the most reactive in  $S_N2$  reactions, followed by primary alkyl halides such as ethyl and propyl. Alkyl branching next to the leaving group, as in isopropyl halides ( $2^\circ$ ), slows the reaction greatly, and further branching, as in *tert*-butyl halides ( $3^\circ$ ), effectively halts the reaction. Even branching one carbon removed from the leaving group, as

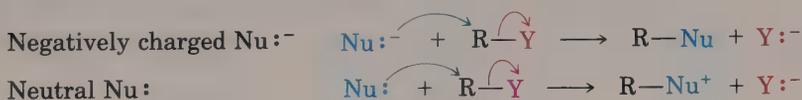
in 2,2-dimethylpropyl (*neopentyl*) halides, greatly slows nucleophilic displacement. Clearly, S<sub>N</sub>2 reactions can occur only at relatively unhindered sites and are normally useful only with methyl halides, primary halides, and a few simple secondary halides.

Although not shown in the preceding reactivity order, vinylic halides (R<sub>2</sub>C=CRX) and aryl halides are completely unreactive toward attempted S<sub>N</sub>2 displacements. This lack of reactivity probably is due to steric factors, since the incoming nucleophile would have to approach in the plane of the carbon–carbon double bond in order to be able to carry out back-side displacements.



## THE ATTACKING NUCLEophile

The nature of the attacking nucleophile is a second variable that has a major effect on the S<sub>N</sub>2 reaction. Any species, either neutral or negatively charged, can act as a nucleophile as long as it has an unshared pair of electrons (that is, as long as it's a Lewis base). If the nucleophile is negatively charged, the product is neutral, but if the nucleophile is neutral, the product is positively charged.

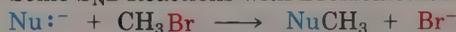


Because of the wide scope of nucleophilic substitution reactions, a great many kinds of products can be prepared from alkyl halides and tosylates. Table 11.1 lists some common nucleophiles and shows the products of their reactions with bromomethane.

Although all the S<sub>N</sub>2 reactions shown in Table 11.1 take place as indicated, some are much faster than others. What are the reasons for the reactivity differences? Why do some reagents appear to be much more “nucleophilic” than others?

The answers to these questions aren't straightforward and aren't yet fully understood. Part of the problem is that the very term *nucleophilic* is imprecise. Although most chemists use the term *nucleophilicity* to mean a measure of the affinity of a species for a carbon atom in the S<sub>N</sub>2 reaction, the reactivity of a given nucleophile can change somewhat from one reaction to the next.

The exact nucleophilicity of a species in a given reaction depends on the nature of the substrate, the identity of the solvent, and even the concentration of the reagents. In order to speak with any precision, we must

TABLE 11.1 Some S<sub>N</sub>2 Reactions with Bromomethane:

Attacking nucleophile		Product	
Formula	Name	Formula	Name
H <sup>-</sup>	Hydride	CH <sub>4</sub>	Methane
CH <sub>3</sub> S <sup>-</sup>	Methanethiolate	CH <sub>3</sub> SCH <sub>3</sub>	Dimethyl sulfide
HS <sup>-</sup>	Hydrosulfide	HSCH <sub>3</sub>	Methane thiol
N≡C <sup>-</sup>	Cyanide	N≡CCH <sub>3</sub>	Acetonitrile
I <sup>-</sup>	Iodide	ICH <sub>3</sub>	Iodomethane
HO <sup>-</sup>	Hydroxide	HOCH <sub>3</sub>	Methanol
CH <sub>3</sub> O <sup>-</sup>	Methoxide	CH <sub>3</sub> OCH <sub>3</sub>	Dimethyl ether
N=N=N <sup>-</sup>	Azide	N <sub>3</sub> CH <sub>3</sub>	Azidomethane
Cl <sup>-</sup>	Chloride	ClCH <sub>3</sub>	Chloromethane
CH <sub>3</sub> CO <sub>2</sub> <sup>-</sup>	Acetate	CH <sub>3</sub> CO <sub>2</sub> CH <sub>3</sub>	Methyl acetate
H <sub>3</sub> N:	Ammonia	H <sub>3</sub> N <sup>+</sup> CH <sub>3</sub> Br <sup>-</sup>	Methylammonium bromide
(CH <sub>3</sub> ) <sub>3</sub> N:	Trimethylamine	(CH <sub>3</sub> ) <sub>3</sub> N <sup>+</sup> CH <sub>3</sub> Br <sup>-</sup>	Tetramethylammonium bromide

study the relative reactivity of various nucleophiles on a *single* substrate in a *single* solvent system. Much work has been carried out on the S<sub>N</sub>2 reactions of bromomethane in aqueous ethanol with the following results:

$$\text{CH}_3\text{Br} + \text{Nu}^- \longrightarrow \text{CH}_3\text{Nu} + \text{Br}^-$$

Nu =	HS <sup>-</sup>	CN <sup>-</sup>	I <sup>-</sup>	CH <sub>3</sub> O <sup>-</sup>	HO <sup>-</sup>	Cl <sup>-</sup>	NH <sub>3</sub>	H <sub>2</sub> O
Relative reactivity	125	125	100	25	16	1.0	0.7	0.001

More reactive ← Reactivity as nucleophile → Less reactive

Although precise explanations for the observed nucleophilicities aren't known, some trends can be detected in the data:

1. Nucleophilicity roughly parallels basicity when comparing nucleophiles that have the same attacking atom. Hydroxide ion, for example, is both more basic and more nucleophilic than water. Since “nucleophilicity” measures the affinity of a Lewis base for a carbon atom in the S<sub>N</sub>2 reaction, and “basicity” measures the affinity of a base for a proton, it's easy to see why there might be a rough correlation between the two kinds of behavior.

2. Nucleophilicity usually increases on going down a column of the periodic table. Thus, HS<sup>-</sup> is more nucleophilic than HO<sup>-</sup>, and the halide reactivity order is I<sup>-</sup> > Br<sup>-</sup> > Cl<sup>-</sup>.

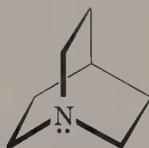
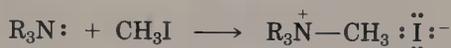
PROBLEM .....

11.4 What products would you expect from reaction of 1-bromobutane with these reagents?

- (a) NaI    (b) KOH    (c) H—C≡C—Li    (d) NH<sub>3</sub>

PROBLEM .....

11.5 The tertiary amine base quinuclidine reacts with CH<sub>3</sub>I 50 times as fast as triethylamine. Can you suggest a reason for this difference?



Quinuclidine



Triethylamine

PROBLEM .....

11.6 Which reagent in each of the following pairs is more nucleophilic? Justify your choices.

- (a) (CH<sub>3</sub>)<sub>2</sub>N<sup>-</sup> and (CH<sub>3</sub>)<sub>2</sub>NH    (b) (CH<sub>3</sub>)<sub>3</sub>B and (CH<sub>3</sub>)<sub>3</sub>N    (c) H<sub>2</sub>O and H<sub>2</sub>S

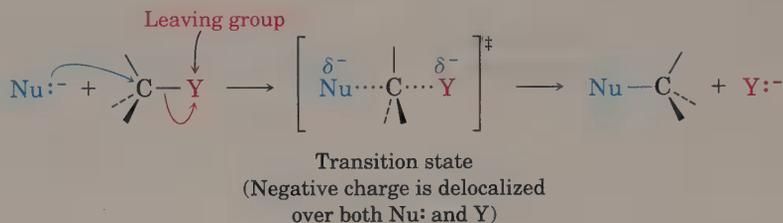
## THE LEAVING GROUP

Another variable that can strongly affect the S<sub>N</sub>2 reaction is the nature of the group displaced by the attacking nucleophile: the **leaving group**. Since the leaving group is expelled with a negative charge in most S<sub>N</sub>2 reactions, we might expect the best leaving groups to be those that best stabilize the negative charge. Furthermore, since the stability of an anion is related to basicity, we can also say that the best leaving groups should be the weakest bases.

As indicated on the following page, the weakest bases (anions derived from the strongest acids) are indeed the best leaving groups. The *p*-toluenesulfonate (tosylate) leaving group is very easily displaced, as are iodide and bromide ion, but chloride and fluoride ion are much less effective as leaving groups.

	TosO <sup>-</sup>	I <sup>-</sup>	Br <sup>-</sup>	Cl <sup>-</sup>	F <sup>-</sup>	HO <sup>-</sup> , H <sub>2</sub> N <sup>-</sup> , RO <sup>-</sup>
Relative reactivity	60	30	10	0.2	0.001	≈ 0
						
	More reactive					Less reactive

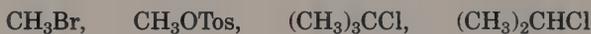
The reason that stable anions (weak bases) make good leaving groups can be understood by looking at the transition state. In the transition state for an S<sub>N</sub>2 reaction, the charge is distributed over both the attacking nucleophile and the leaving group. The greater the extent of charge stabilization by the leaving group, the more stable the transition state and the more rapid the reaction.



It's just as important to know which are *poor* leaving groups as to know which are good, and the preceding data clearly indicate that F<sup>-</sup>, HO<sup>-</sup>, RO<sup>-</sup>, and H<sub>2</sub>N<sup>-</sup> are not displaced by nucleophiles. In other words, alkyl fluorides, alcohols, ethers, and amines do not undergo S<sub>N</sub>2 reactions under normal circumstances.

#### PROBLEM .....

- 11.7 Rank the following compounds in order of their expected reactivity toward S<sub>N</sub>2 reaction:

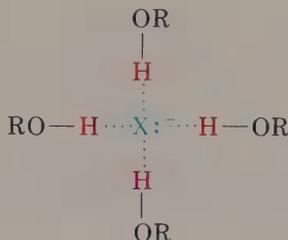


### THE SOLVENT

The rates of many S<sub>N</sub>2 reactions are affected by the solvent. Protic solvents, which contain -OH groups, are generally the worst solvents for S<sub>N</sub>2 reactions; **polar aprotic solvents**, which have strong dipoles but don't have -OH or -NH groups, are the best.

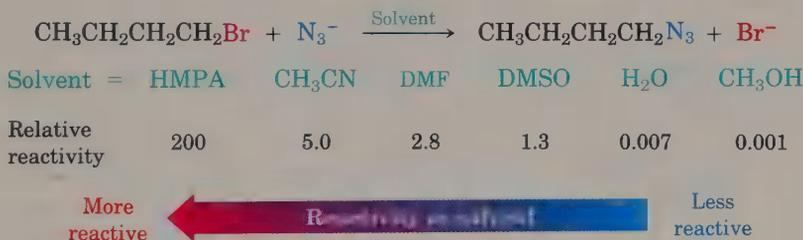
Protic solvents such as methanol and ethanol slow down S<sub>N</sub>2 reactions by affecting the energy level of the nucleophilic reactant rather than the energy level of the transition state. Protic solvent molecules are able to form hydrogen bonds to negatively charged nucleophiles, orienting themselves

into a “cage” around the nucleophile. This **solvation** strongly stabilizes the nucleophile, decreasing its reactivity toward electrophiles in the S<sub>N</sub>2 reaction.



A solvated anion  
(reduced nucleophilicity due to enhanced ground-state stability)

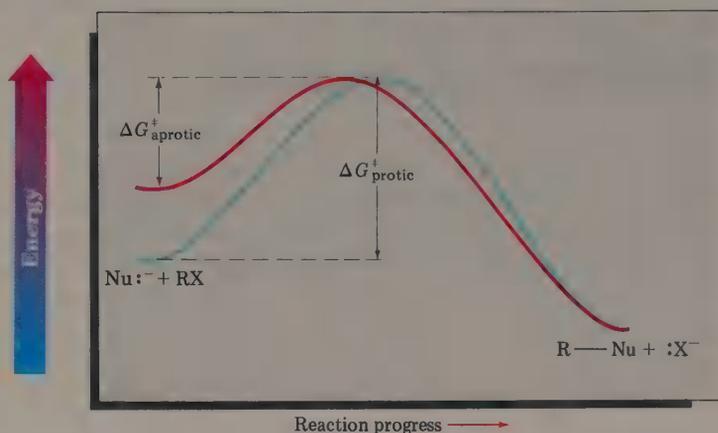
In contrast to protic solvents, polar aprotic solvents favor S<sub>N</sub>2 reactions. Particularly valuable are acetonitrile, CH<sub>3</sub>CN; dimethylformamide, (CH<sub>3</sub>)<sub>2</sub>NCHO (abbreviated DMF); dimethyl sulfoxide, (CH<sub>3</sub>)<sub>2</sub>SO (abbreviated DMSO); and hexamethylphosphoramide, [(CH<sub>3</sub>)<sub>2</sub>N]<sub>3</sub>PO (abbreviated HMPA). These solvents are able to dissolve many salts because of their high polarity, but they tend to surround the metal cations rather than the nucleophilic anions. As a result, the unsolvated anions have a greater effective nucleophilicity in these solvents, and S<sub>N</sub>2 reactions take place at correspondingly faster rates, as indicated below. Rate increases of over a million-fold have been observed in going from methanol to hexamethylphosphoramide for the reaction of azide ion with 1-bromobutane:



Polar aprotic solvents increase the rate of S<sub>N</sub>2 reactions by affecting the energy level of the nucleophilic reactant rather than the energy level of the transition state. They lower ΔG<sup>‡</sup> by destabilizing (raising) the ground-state energy level of the nucleophile (Figure 11.8).

**PROBLEM** .....

- 11.8 Normal organic solvents such as benzene, ether, and chloroform are neither protic nor strongly polar. What effect would you expect these solvents to have on S<sub>N</sub>2 reactions?
- .....

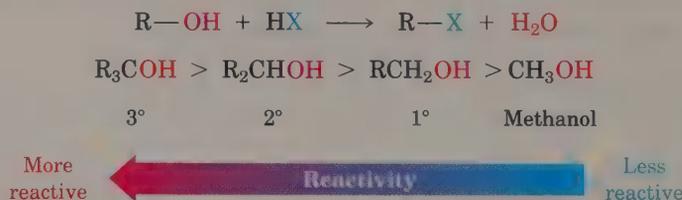


**Figure 11.8** The effect of solvent on  $S_N2$  reactions. The nucleophile is stabilized and less reactive in protic solvents but destabilized and more reactive in polar aprotic solvents.

## 11.6 The $S_N1$ Reaction

As we've seen, the  $S_N2$  reaction is favored by basic nucleophiles like hydroxide ion and is disfavored by protic solvents such as water and alcohol. As we've also seen, the reaction is sensitive to steric factors: Primary substrates react rapidly, secondary substrates react more slowly, and tertiary substrates are essentially inert to back-side attack by nucleophiles.

Remarkably, a completely different picture emerges when different reaction conditions are used. When treated in protic solvents with nonbasic nucleophiles under neutral or acidic conditions, tertiary substrates often react several thousand times *faster* than primary or secondary ones. We noted in Section 10.8, for example, that the substitution reaction of alcohols with  $HX$  to yield alkyl halides is fastest for tertiary alcohols and slowest for methanol:



The same trend is noted in many substitution reactions in which substrates are heated with nonbasic nucleophiles in protic solvents: Tertiary substrates react much faster than primary or secondary ones. For example,

the relative rates of reaction of some alkyl halides with water shown below indicate that 2-bromo-2-methylpropane is more than 1 million times as reactive as bromoethane.

	$\text{RBr} + \text{H}_2\text{O} \longrightarrow \text{ROH} + \text{HBr}$			
	$(\text{CH}_3)_3\text{CBr}$	$(\text{CH}_3)_2\text{CHBr}$	$\text{CH}_3\text{CH}_2\text{Br}$	$\text{CH}_3\text{Br}$
Relative reactivity	1,200	0.012	0.001	0.001
More reactive				Less reactive

What's going on here? Clearly, nucleophilic substitution reactions are occurring, yet the reactivity order seems backward. These reactions can't be taking place by the mechanism we've been discussing, and we must therefore conclude that *an alternative substitution mechanism exists*. This alternative mechanism is called the **S<sub>N</sub>1 reaction** (for *substitution, nucleophilic, unimolecular*). Let's see what evidence is available concerning the S<sub>N</sub>1 reaction.

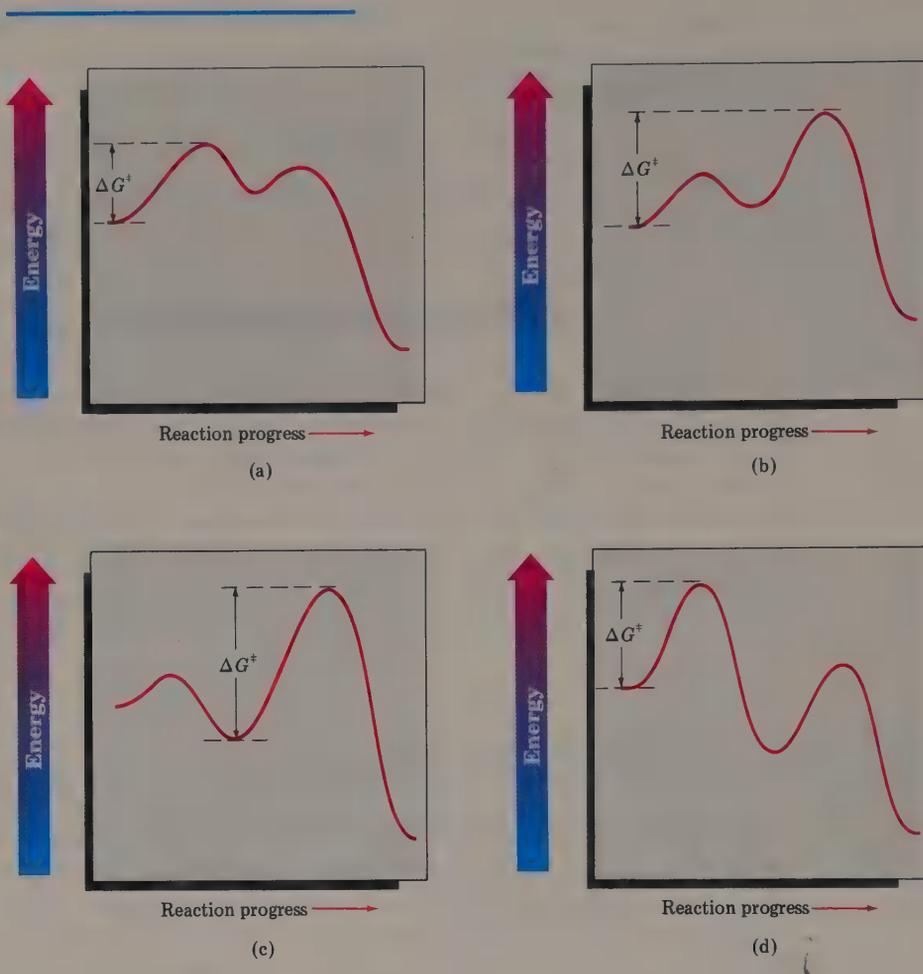
## 11.7 Kinetics of the S<sub>N</sub>1 Reaction

The reaction of water with 2-bromo-2-methylpropane seems analogous to the reaction of hydroxide ion with bromomethane, and we might therefore expect to observe second-order kinetics. In fact, we do not. We find instead that the reaction rate is dependent only on the alkyl halide concentration but is independent of water concentration. In other words, the reaction is a **first-order process**. Only one molecule is involved in the step whose kinetics are measured. We can write the rate expression as follows:

$$\begin{aligned} \text{Reaction rate} &= \text{Rate of disappearance of alkyl halide} \\ &= k \times [\text{RX}] \end{aligned}$$

The rate of this S<sub>N</sub>1 reaction is equal to a rate coefficient  $k$  times the alkyl halide concentration. *Nucleophile concentration does not appear in the rate expression*. How can this result be explained? To answer this question, we must first learn more about kinetics measurements.

Many organic reactions are rather complex and occur in successive steps. One of these steps is usually slower than the others, and we call this the **rate-limiting step**. No reaction can proceed faster than its rate-limiting step, which acts as a kind of traffic jam or bottleneck. The overall reaction rate that we actually measure in a kinetics experiment is determined by the height of the highest *cumulative* energy barrier ( $\Delta G^\ddagger$ ) between a low point and a subsequent high point in the energy diagram of the reaction. The hypothetical reaction energy diagrams in Figure 11.9 illustrate the idea of the rate-limiting step.

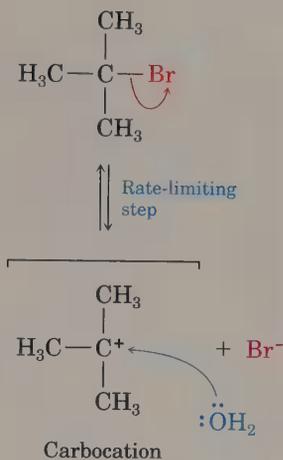


**Figure 11.9** Hypothetical reaction energy diagrams. The rate-limiting step in each is determined by the greatest cumulative difference in height between a low point and a subsequent high point.

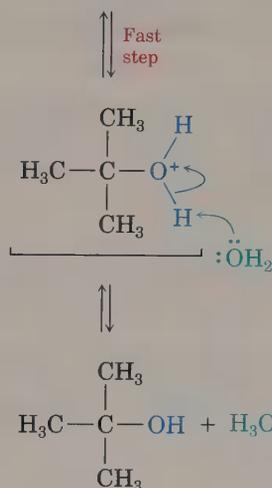
In Figure 11.9(a), the rate-limiting step is simply the height of the barrier from the starting material to the first transition state, whereas in (b) the rate-limiting step represents the *cumulative* barrier from starting material to the highest-energy, *second* transition state. In (c) the step corresponding to the highest energy barrier is that from the stable intermediate to the second transition state, and in (d) the highest barrier is that from starting material to the first transition state. With this as background, let's look further into the kinetics of the  $S_N1$  reaction.

The observation of first-order kinetics for the  $S_N1$  reaction of 2-bromo-2-methylpropane with water indicates that the alkyl halide is involved in a **unimolecular** rate-limiting step. In other words, 2-bromo-2-methylpropane undergoes some manner of spontaneous reaction without assistance

Spontaneous dissociation of the alkyl bromide occurs in a slow, rate-limiting step to generate a carbocation intermediate plus bromide ion.



The carbocation intermediate reacts with water as nucleophile in a fast step to yield protonated alcohol as product.

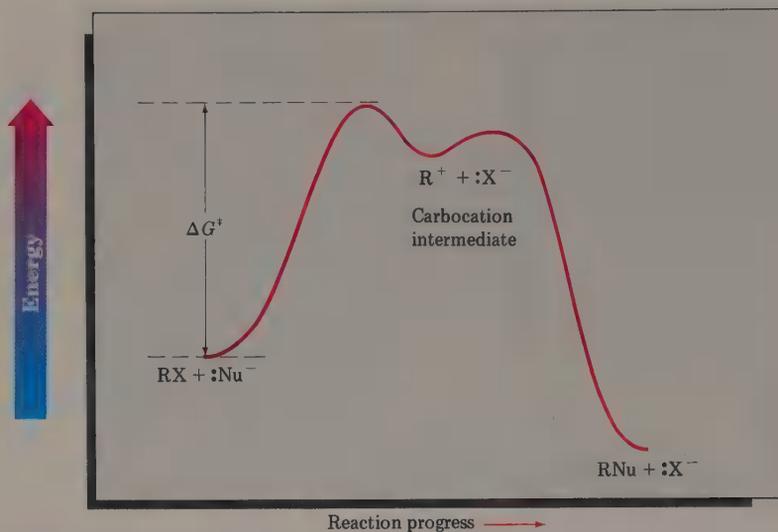


Loss of proton from the protonated alcohol intermediate then gives the neutral alcohol product.

**Figure 11.10** The mechanism of the  $S_N1$  reaction of 2-bromo-2-methylpropane with water. Three steps are involved, with the first one rate-limiting.

from the nucleophile. The mechanism shown in Figure 11.10 accounts for the kinetic observations.

Unlike what happens in an  $S_N2$  reaction, in which the leaving group is displaced at the same time that the incoming nucleophile is approaching, an  $S_N1$  reaction takes place by unassisted loss of the leaving group *before* the incoming nucleophile approaches. 2-Bromo-2-methylpropane spontaneously dissociates to the *tert*-butyl carbocation plus bromide ion in a slow, rate-limiting step, and the intermediate ion is then immediately trapped by nucleophilic water in a fast step so that first-order kinetics are obtained. *Water plays no role in the step whose rate is measured by kinetics.* The reaction energy diagram is shown in Figure 11.11.

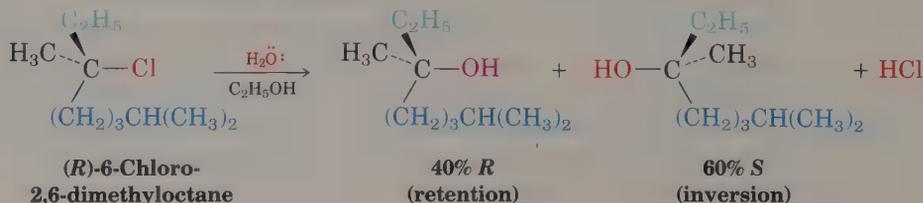


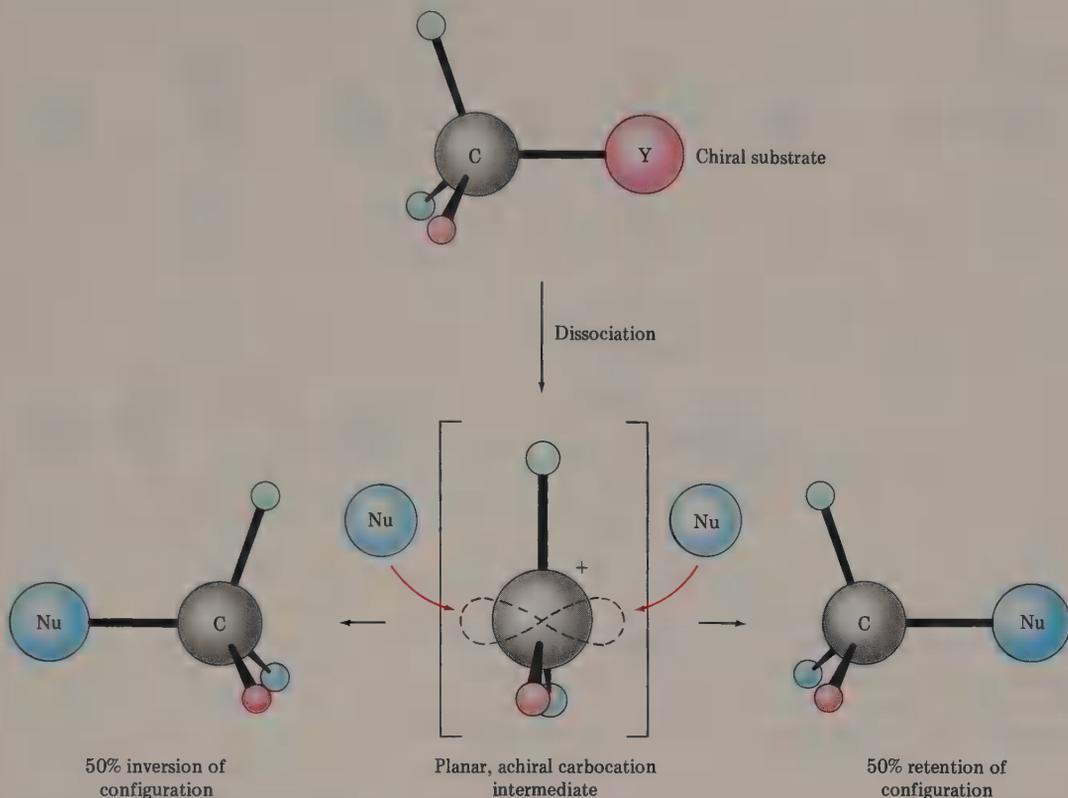
**Figure 11.11** A reaction energy diagram for an  $S_N1$  reaction. The rate-limiting step is the spontaneous dissociation of the alkyl halide.

## 11.8 Stereochemistry of the $S_N1$ Reaction

If  $S_N1$  reactions occur through carbocation intermediates, the stereochemical consequences should differ from those for  $S_N2$  reactions. Since carbocations are planar,  $sp^2$ -hybridized species, they are achiral. Thus, if we carry out an  $S_N1$  reaction on one enantiomer of a *chiral* starting material and go through an *achiral* carbocation intermediate, then the product must be optically inactive. The symmetrical intermediate carbocation can be attacked by a nucleophile equally well from either side, leading to a 50:50 mixture of enantiomers—a racemic mixture (Figure 11.12).

The prediction that  $S_N1$  reactions on optically active substrates should lead to racemic products is exactly what's observed. Surprisingly, though, few  $S_N1$  displacements occur with *complete* racemization. Most give a minor (0–20%) excess of inversion. For example, the reaction of optically active (*R*)-6-chloro-2,6-dimethyloctane with water leads to an alcohol product that is approximately 80% racemized and 20% inverted (80% *R,S* + 20% *S* is the same as 40% *R* + 60% *S*):





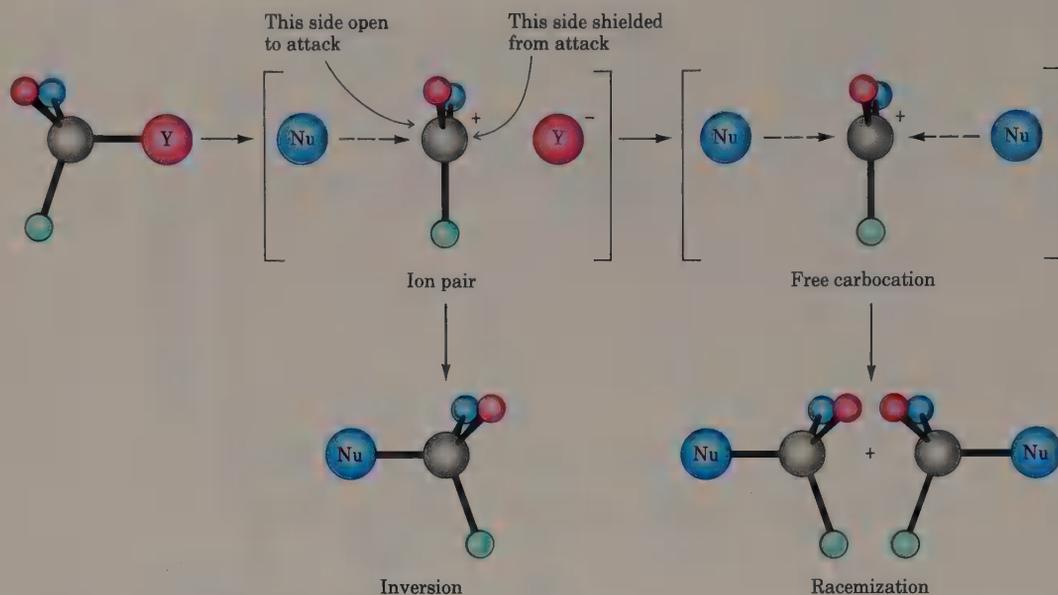
**Figure 11.12** Stereochemistry of the  $S_N1$  reaction. An optically active starting material must give a racemic product.

The situation is complex, and the reasons for the lack of complete racemization in most  $S_N1$  reactions aren't completely clear. An attractive suggestion, first proposed by Saul Winstein,<sup>4</sup> is that **ion pairs** are involved. According to this suggestion, dissociation of the substrate occurs to give a complex in which the two ions are still loosely associated and in which the carbocation is effectively shielded from nucleophilic attack on one side by the departing ion. If a certain amount of substitution occurs before the two ions fully diffuse away from each other, then a net inversion of configuration will be observed (Figure 11.13).

**PROBLEM** .....

- 11.9** What product(s) would you expect from reaction of (*S*)-3-chloro-3-methyloctane with acetic acid? Show the stereochemistry of both starting material and product.

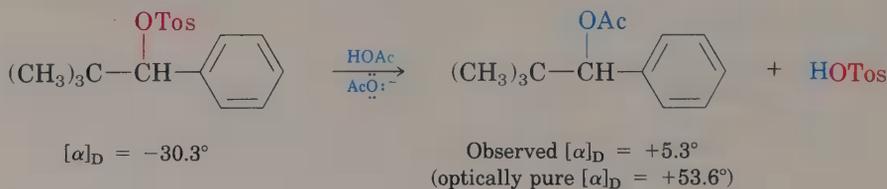
<sup>4</sup>Saul Winstein (1912–1969); b. Montreal; Ph.D., California Institute of Technology (Lucas); professor, University of California, Los Angeles (1942–1969).



**Figure 11.13** The ion-pair hypothesis. The leaving group effectively shields one side of the developing carbocation intermediate from attack by added nucleophile, thereby leading to some inversion of configuration rather than complete racemization.

**PROBLEM** .....

- 11.10** Among the numerous examples of  $S_N1$  reactions that occur with incomplete racemization is one reported by Winstein in 1952. The optically pure tosylate of 2,2-dimethyl-1-phenyl-1-propanol ( $[\alpha]_D = -30.3^\circ$ ) was heated in acetic acid to yield the corresponding acetate ( $[\alpha]_D = +5.3^\circ$ ). If complete inversion had occurred, the optically pure acetate would have had  $[\alpha]_D = +53.6^\circ$ . What percentage racemization and what percentage inversion occurred in this reaction?



## 11.9 Characteristics of the $S_N1$ Reaction

Just as the  $S_N2$  reaction is strongly influenced by such variables as solvent, leaving group, substrate structure, and nature of the attacking nucleophile, the  $S_N1$  reaction is similarly influenced. Factors that lower  $\Delta G^\ddagger$ , either by

stabilizing the transition state leading to carbocation formation or by raising the reactant energy level, favor faster S<sub>N</sub>1 reactions. Conversely, factors that raise ΔG<sup>‡</sup>, either by destabilizing the transition state leading to a carbocation or by lowering reactant energy level, slow down the S<sub>N</sub>1 reaction.

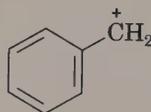
## THE SUBSTRATE

According to the Hammond postulate, any factor that stabilizes a high-energy intermediate should also stabilize the transition state leading to that intermediate. Since the rate-limiting step in the S<sub>N</sub>1 reaction is the spontaneous, unimolecular dissociation of the substrate, we would expect the reaction to be favored whenever stabilized carbocation intermediates are formed. This is exactly what's found: *The more stable the carbocation intermediate, the faster the S<sub>N</sub>1 reaction.*

We've already seen (Section 6.11) that the stability order of alkyl carbocations is 3° > 2° > 1° > -CH<sub>3</sub>. To this list we must now add the resonance-stabilized allyl and benzyl cations:



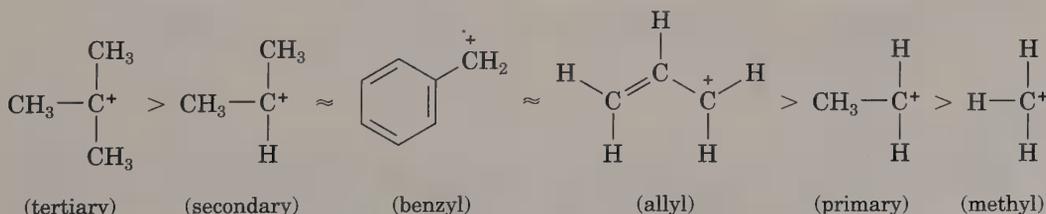
Allyl carbocation



Benzyl carbocation

We saw in Sections 10.6 and 10.7 that allylic radicals have unusual stability because the unpaired electron can be delocalized over an extended pi orbital system. The same is true for allylic and benzylic carbocations (**benzylic** = next to aromatic ring). Delocalization of the positive charge over the extended pi orbital system of these carbocations results in unusual stability. As Figure 11.14 indicates, an allylic cation has two equivalent resonance forms. In one form the double bond is on the "left," and in the other form the double bond is on the "right." A benzylic cation, however, has *four* resonance forms. These four forms are not equivalent, but all four make substantial contributions to the overall resonance hybrid.

As a result of the resonance stabilization in allylic and benzylic carbocations, carbocations have the following stability order. Note that a *primary* allylic or benzylic carbocation is approximately as stable as a *secondary* alkyl carbocation. Similarly, a *secondary* allylic or benzylic carbocation is about as stable as a *tertiary* alkyl carbocation.

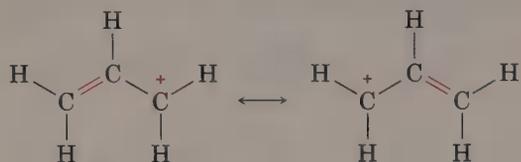


More  
stable

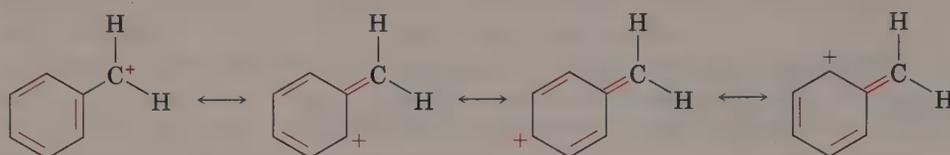
Stability

Less  
stable

Allyl carbocation



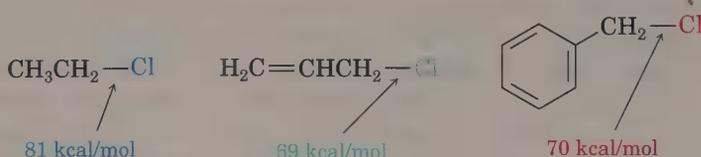
Benzylic carbocation



**Figure 11.14** Resonance forms of allylic and benzylic carbocations.

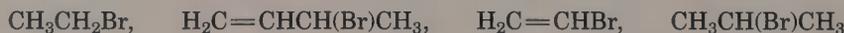
The stability order of carbocations is exactly the order of  $S_N1$  reactivity for alkyl halides and tosylates.

Parenthetically, it should be noted that allylic and benzylic substrates are particularly reactive in  $S_N2$  reactions as well as in  $S_N1$  reactions. Allylic and benzylic C–X bonds are weaker by about 10 kcal/mol than are the corresponding saturated bonds and are therefore more easily broken.



**PROBLEM** .....

11.11 Rank the following alkyl halides in order of their expected  $S_N1$  reactivity.



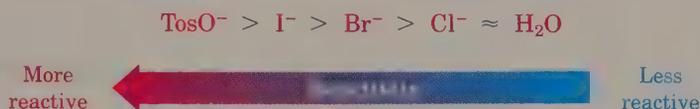
**PROBLEM** .....

11.12 How can you account for the fact that 3-bromo-1-butene and 1-bromo-2-butene undergo  $S_N1$  reaction at nearly the same rate even though one is a secondary halide and the other is primary?

.....

## THE LEAVING GROUP

During the discussion of S<sub>N</sub>2 reactivity, we reasoned that the best leaving groups should be those that are most stable—that is, the conjugate bases of strong acids. An identical reactivity order is found for the S<sub>N</sub>1 reaction, because the leaving group is intimately involved in the rate-limiting step. Thus, we find the S<sub>N</sub>1 reactivity order to be

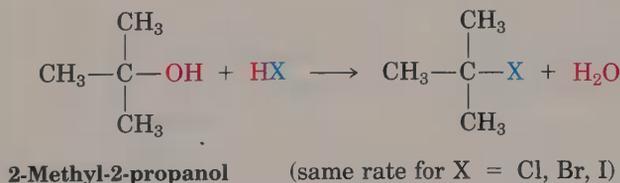


Note that in the S<sub>N</sub>1 reaction, which is often carried out under acidic conditions, neutral water can act as a leaving group. This is the case, for example, when an alkyl halide is prepared from a tertiary alcohol by reaction with HBr or HCl (Section 10.8). The alcohol is first protonated and then loses water to generate a carbocation. Reaction of the carbocation with halide ion then yields the alkyl halide (Figure 11.15).

Realizing that an S<sub>N</sub>1 reaction is involved in the conversion of alcohols to alkyl halides makes it easier to understand why the reaction works well only for tertiary alcohols: Tertiary alcohols react fastest because they give the most stable carbocation intermediates.

## THE NUCLEOPHILE

The nature of the attacking nucleophile plays a major role in the S<sub>N</sub>2 reaction. Should it play an equally major role in the S<sub>N</sub>1 reaction? The answer is no. The S<sub>N</sub>1 reaction, by its very nature, occurs through a rate-limiting step in which the added nucleophile plays no kinetic role. The nucleophile does not enter into the reaction until after rate-limiting dissociation has occurred and thus cannot affect the rate of the reaction. The reaction of 2-methyl-2-propanol with HX, for example, occurs at the same rate regardless of whether X is Cl, Br, or I:



## PROBLEM .....

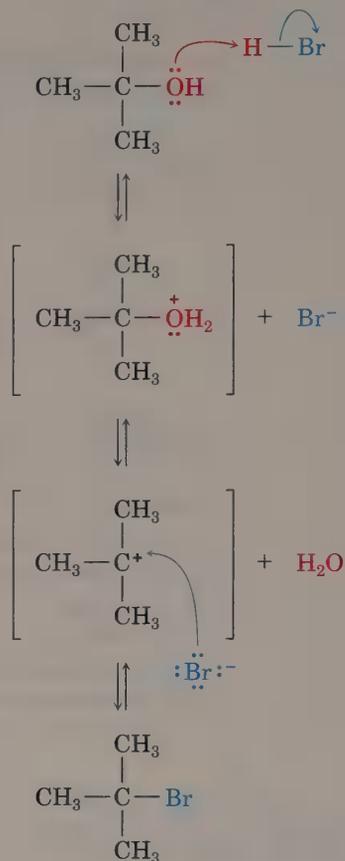
- 11.13 How do you account for the fact that 1-chloro-1,2-diphenylethane reacts with the nucleophiles fluoride ion and triethylamine at exactly the same rate?
- .....

The  $\text{-OH}$  group is first protonated by  $\text{HBr}$ .

Spontaneous dissociation of the protonated alcohol occurs in a slow, rate-limiting step to yield a carbocation intermediate plus water.

The carbocation intermediate reacts with bromide ion in a fast step to yield the neutral substitution product.

Carbocation



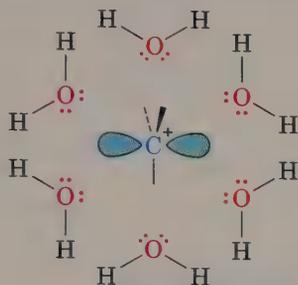
**Figure 11.15** The mechanism of the  $\text{S}_{\text{N}}1$  reaction of a tertiary alcohol with  $\text{HBr}$  to yield an alkyl halide. Neutral water is the leaving group.

## THE SOLVENT

What about solvent? Does it have the same kind of effect in  $\text{S}_{\text{N}}1$  reactions that it has in  $\text{S}_{\text{N}}2$  reactions? The answer is both yes and no. Yes, solvents have a large effect on  $\text{S}_{\text{N}}1$  reactions just as they do on  $\text{S}_{\text{N}}2$  reactions, but no, the reasons for their effects are not the same. Solvent effects in the  $\text{S}_{\text{N}}2$  reaction are due to stabilization or destabilization of the nucleophile *reactant*. Solvent effects in the  $\text{S}_{\text{N}}1$  reaction, however, are due to stabilization of the *transition state*. Let's look again at the Hammond postulate to see the effect of solvent on the transition state.

The Hammond postulate (Section 6.12) says that any factor stabilizing the intermediate carbocation should increase the rate of reaction. Solvation—the stabilization of the carbocation by interaction with solvent molecules—can have just such an effect. Although the exact nature of carbocation stabilization by solvent is not easily defined, we might picture

the solvent molecules orienting themselves around the cation in such a manner that the electron-rich ends of the solvent dipoles face the positive charge (Figure 11.16).



**Figure 11.16** Solvation of a carbocation by water. The electron-rich oxygen atoms of solvent surround the positively charged carbocation to stabilize it.

The properties of a solvent that contribute to its ability to stabilize ions by solvation aren't fully understood but are related to the polarity of the solvent. Polar solvents such as water, methanol, dimethyl sulfoxide, and so forth are good at solvating ions, but most ether and hydrocarbon solvents are very poor at solvating ions.

Solvent polarity is usually expressed in terms of **dielectric constants**,  $\epsilon$ , which measure the ability of a solvent to act as an insulator of electric charges. Solvents of low dielectric constant such as hydrocarbons are nonpolar, whereas solvents of high dielectric constant such as water are polar. Table 11.2 lists the dielectric constants of some common solvents.

**TABLE 11.2** Dielectric Constants of Some Common Solvents

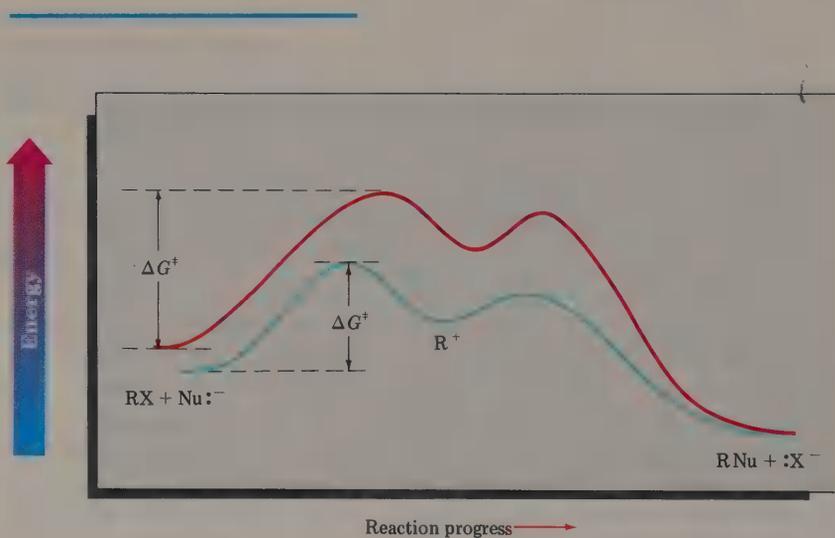
Name	Dielectric constant	Name	Dielectric constant
<i>Aprotic solvents</i>		<i>Protic solvents</i>	
Hexane	1.9	Acetic acid	6.2
Benzene	2.3	Ethanol	24.3
Diethyl ether	4.3	Methanol	33.6
Chloroform	4.8	Formic acid	58.0
Hexamethylphosphoramide (HMPA)	30	Water	80.4
Dimethylformamide (DMF)	38		
Dimethyl sulfoxide (DMSO)	48		

$S_N1$  reactions take place much more rapidly in highly polar solvents than in nonpolar solvents. In the reaction of 2-chloro-2-methylpropane, for example, a rate increase of 100,000 is observed on going from the polar solvent ethanol to the even more polar solvent water. The rate increases on going from hydrocarbon solvents to water are too large to measure accurately.



	Water	80% aqueous ethanol	40% aqueous ethanol	Ethanol
Relative reactivity	100	14	0.1	0.001
More reactive	←			Less reactive

It should be emphasized again that both  $S_N1$  and  $S_N2$  reactions show large solvent effects, but that they do so for different reasons.  $S_N2$  reactions are disfavored by protic solvents because the *ground-state energy level* of the attacking nucleophile is lowered by solvation.  $S_N1$  reactions, however, are favored by protic solvents because the *transition-state energy level* leading to carbocation intermediate is lowered by solvation. Compare the  $S_N1$  reaction energy diagram in Figure 11.17 to that in Figure 11.8, where the effect of solvent on the  $S_N2$  reaction is illustrated.



**Figure 11.17** The effect of solvent on  $S_N1$  reactions. Both ground-state and transition-state energy levels are lowered by solvation in polar solvents, but the effect on transition state is much greater. (Red curve, nonpolar solvent; green curve, polar solvent.)

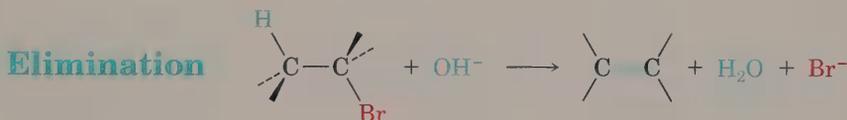
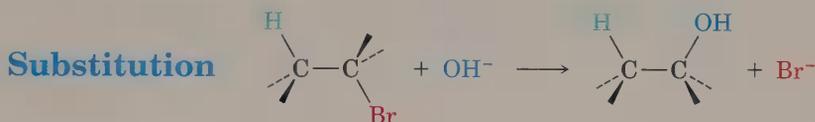
## PROBLEM .....

- 11.14 As indicated in Problem 11.3, halides such as the one shown here are inert to  $S_N2$  displacement. Perhaps more surprisingly, they're also unreactive to  $S_N1$  substitution, even though they're tertiary. Considering the planar geometry of carbocations, can you suggest a reason for this low reactivity? (Molecular models may be helpful.)



## 11.10 Elimination Reactions of Alkyl Halides

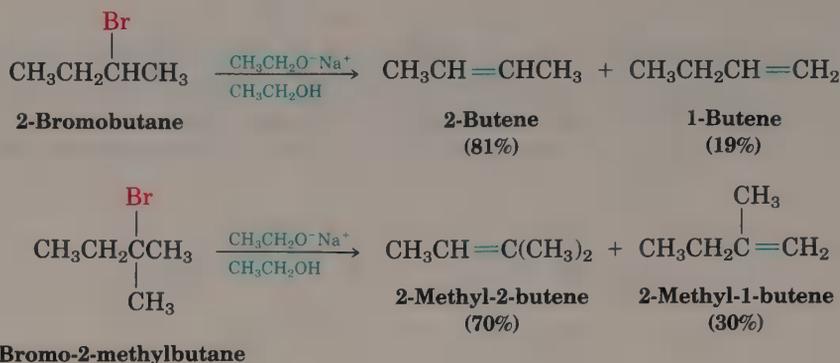
We began this chapter by saying that two kinds of reactions are possible when a nucleophile/Lewis base attacks an alkyl halide. Either the reagent can attack at carbon and substitute for the halide or it can attack at hydrogen and cause elimination of HX to form an alkene.



Elimination reactions are more complex than substitution reactions for several reasons. There is, for example, the problem of regiochemistry: What products result from dehydrohalogenation of unsymmetrical halides? In fact, elimination reactions almost always give *mixtures* of alkene products, and the best we can usually do is to predict which will be the major product.

According to a rule formulated by the Russian chemist Alexander Zaitsev,<sup>5</sup> base-induced elimination reactions generally give the more highly substituted alkene product—that is, the alkene with more alkyl substituents on the double-bond carbons. In the following two examples, **Zaitsev's rule** is clearly applicable. The more highly substituted alkene product predominates in both cases when sodium ethoxide in ethanol is used as the base.

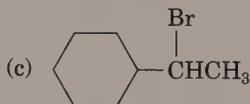
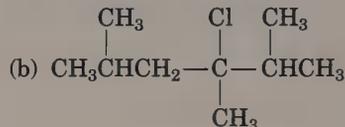
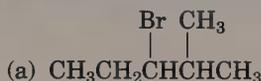
<sup>5</sup>Alexander M. Zaitsev (1841–1910); b. Kasan, Russia (name also spelled *Saytzeff*, according to the German pronunciation).



The elimination of HX from an alkyl halide is a general reaction and an excellent method for preparing alkenes. The subject is complex, however, because elimination reactions can take place through a variety of different mechanistic pathways just as substitutions can. We'll consider two of them: the E1 and E2 reactions.

PROBLEM.....

- 11.15 What products would you expect from elimination reactions of these alkyl halides? Which product will be major in each case?

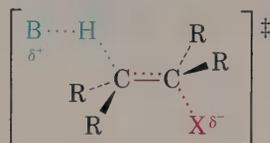
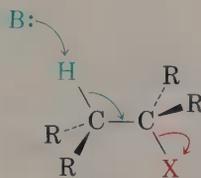


## 11.11 The E2 Reaction

The **E2 reaction** (for *elimination, bimolecular*) takes place when an alkyl halide is treated with a strong base such as hydroxide ion or alkoxide ion ( $\text{RO}^-$ ). It is the most commonly occurring pathway for elimination and can be formulated as shown in Figure 11.18.

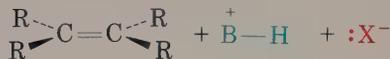
Like the  $\text{S}_{\text{N}}2$  reaction, the E2 reaction takes place in one step without intermediates. As the attacking base begins to abstract a proton from a carbon next to the leaving group, the C–H bond begins to break, a new carbon–carbon double bond begins to form, and the leaving group begins to depart, taking with it the electron pair from the C–X bond.

Base (B:) attacks a neighboring hydrogen and begins to remove the H at the same time as the alkene double bond starts to form and the X group starts to leave.



Transition state

Neutral alkene is produced when the C-H bond is fully broken and the X group has departed with the C-X bond electron pair.

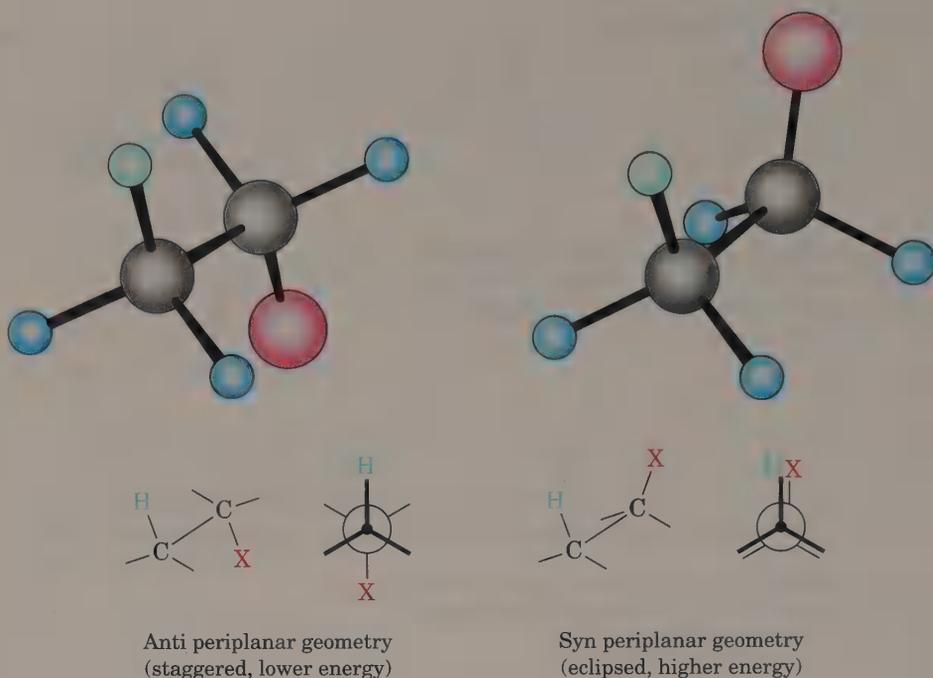


**Figure 11.18** Mechanism of the E2 reaction of an alkyl halide. The reaction takes place in a single step through a transition state in which the double bond begins to form at the same time the H and X groups are leaving.

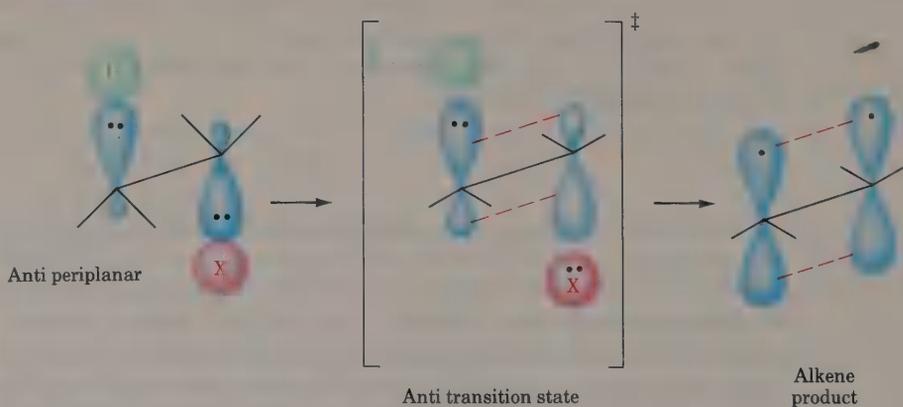
One piece of evidence supporting this mechanism is the measurement of reaction kinetics. Since both base and alkyl halide enter into the rate-limiting step, E2 reactions show second-order kinetics. In other words, E2 reactions follow the rate law:

$$\text{Rate} = k \times [\text{RX}] \times [\text{Base}]$$

A second and more compelling piece of evidence involves the stereochemistry of E2 eliminations. As indicated by a large amount of experimental data, E2 reactions are stereospecific. Elimination always occurs from a **periplanar** geometry, meaning that all four reacting atoms—the hydrogen, the two carbons, and the leaving group—lie in the same plane. Two such geometries are possible: **syn periplanar** geometry, in which the H and the X are on the same side of the molecule; and **anti periplanar** geometry, in which the H and the X are on opposite sides of the molecule. Of the two choices, anti periplanar geometry is energetically preferred because it allows the two carbon centers to adopt a staggered relationship, whereas syn geometry requires that the substituents on carbon be eclipsed.



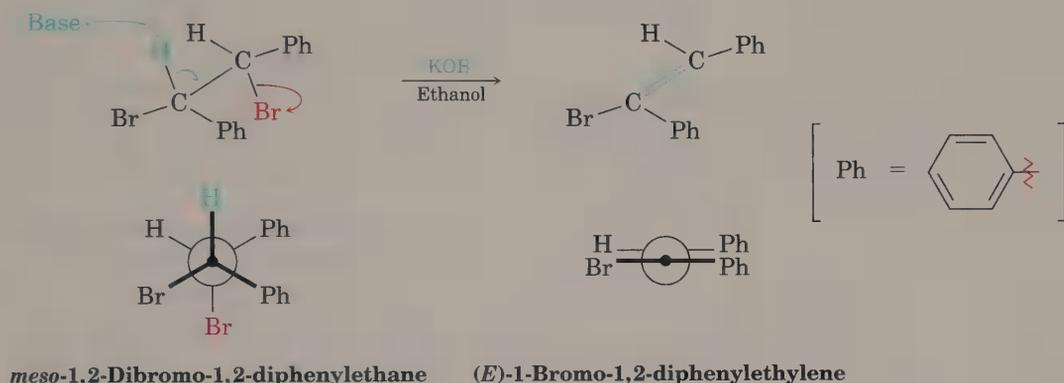
What's so special about periplanar geometry? Because the original C-H and C-X  $sp^3$  sigma orbitals in the starting material must overlap and become  $p$  pi orbitals in the alkene product, there must also be partial overlap in the transition state. This can occur only if all the orbitals are in the same plane to begin with—that is, if they're periplanar (Figure 11.19).



**Figure 11.19** The transition state for the E2 reaction of an alkyl halide with base. Partial overlap of the developing  $p$  orbitals in the transition state requires periplanar geometry of the starting material.

It might help to think of E2 elimination reactions with periplanar geometry as being similar to  $S_N2$  reactions with  $180^\circ$  geometry. In an  $S_N2$  reaction, an electron pair from the incoming nucleophile pushes out the leaving group on the opposite side of the molecule (back-side attack). In an E2 reaction, an electron pair from a neighboring C–H bond pushes out the leaving group on the opposite side of the molecule (anti periplanar).

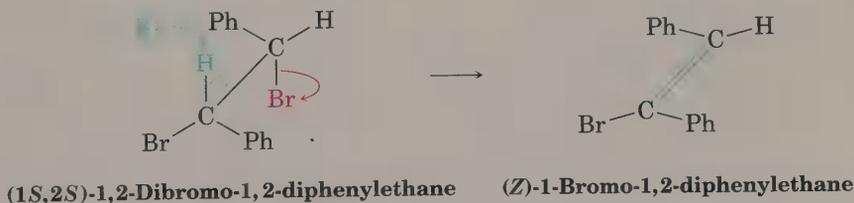
Anti periplanar geometry for E2 eliminations has definite stereochemical consequences that provide strong evidence for the proposed mechanism. To take just one example, *meso*-1,2-dibromo-1,2-diphenylethane undergoes E2 elimination on treatment with base to give only the pure *E* alkene with the benzene rings (phenyl groups) *cis*. None of the isomeric *Z* alkene is formed because the transition state leading to the *Z* alkene would have to have syn periplanar geometry.



### PRACTICE PROBLEM.....

What stereochemistry would you expect for the alkene obtained by E2 elimination of (1*S*,2*S*)-1,2-dibromo-1,2-diphenylethane?

**Solution** First draw (1*S*,2*S*)-1,2-dibromo-1,2-diphenylethane so that you can see its stereochemistry and so that the –H and –Br groups to be eliminated are anti periplanar (molecular models are extremely helpful here). Keeping all substituents in approximately their same positions, eliminate HBr and see what alkene results. The product is (*Z*)-1-bromo-1,2-diphenylethylene.



### PROBLEM.....

**11.16** What stereochemistry would you expect for the alkene obtained by E2 elimination of (1*R*,2*R*)-1,2-dibromo-1,2-diphenylethane? Draw a Newman projection of the reacting conformation.

## PROBLEM .....

- 11.17 1-Chloro-1,2-diphenylethane can undergo E2 elimination to give either *cis*- or *trans*-1,2-diphenylethylene (stilbene). Draw Newman projections of the reactive conformations leading to both *cis*- and *trans*-1,2-diphenylethylene. Look at both conformations and suggest a reason why the *trans* alkene is the major product.



## 11.12 Elimination Reactions and Cyclohexane Conformation

Anti periplanar geometry for E2 reactions is particularly important in cyclohexane rings where rigid chair-like geometry forces a specific relationship between the substituents on neighboring carbon atoms (Section 4.10). As pointed out by Derek Barton<sup>6</sup> in a landmark 1950 paper, much of the chemical reactivity of substituted cyclohexanes is controlled by conformational effects. Let's look at the E2 dehydrohalogenation of chlorocyclohexanes to see the effect of conformation on reactivity.

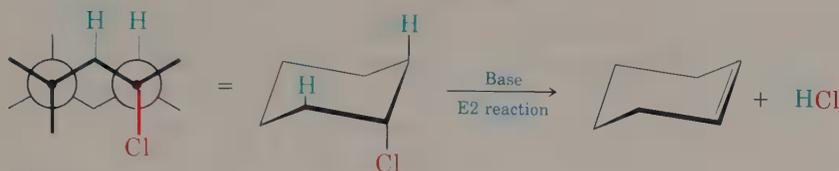
The anti periplanar requirement for E2 reactions can be met in cyclohexanes only if the hydrogen and the halogen are *trans* diaxial (Figure 11.20). If either the halogen or the hydrogen is equatorial, E2 elimination can't occur.

The elimination of HCl from the menthyl chlorides (Figure 11.21) provides a good illustration of this *trans*-diaxial requirement. Neomenthyl chloride undergoes elimination of HCl on reaction with ethoxide ion 200 times as fast as menthyl chloride. Furthermore, neomenthyl chloride yields 3-menthene as the sole alkene product, whereas menthyl chloride yields primarily 2-menthene.

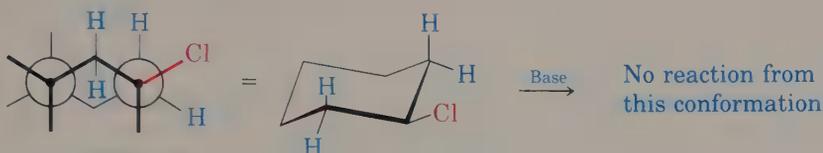
We can understand this difference in reactivity by looking at the most favorable chair conformations of the reactant molecules. Neomenthyl chloride should have the conformation shown in Figure 11.21(a) with the two large substituents, methyl and isopropyl, equatorial. In this conformation, chlorine is rigidly held in an axial orientation—the perfect geometry for E2 elimination with loss of the hydrogen atom at C4 to yield the more substituted alkene product 3-menthene.

<sup>6</sup>Derek H. R. Barton (1918– ); b. Gravesend, England; Ph.D. and D.Sc. London (Heilbron, E. R. H. Jones); professor, Birkbeck College, Harvard, Glasgow, Imperial College, London, Institut de Chimie des Substances Naturelles, Gif-sur-Yvette, France, Texas A.&M.; Nobel Prize (1969).

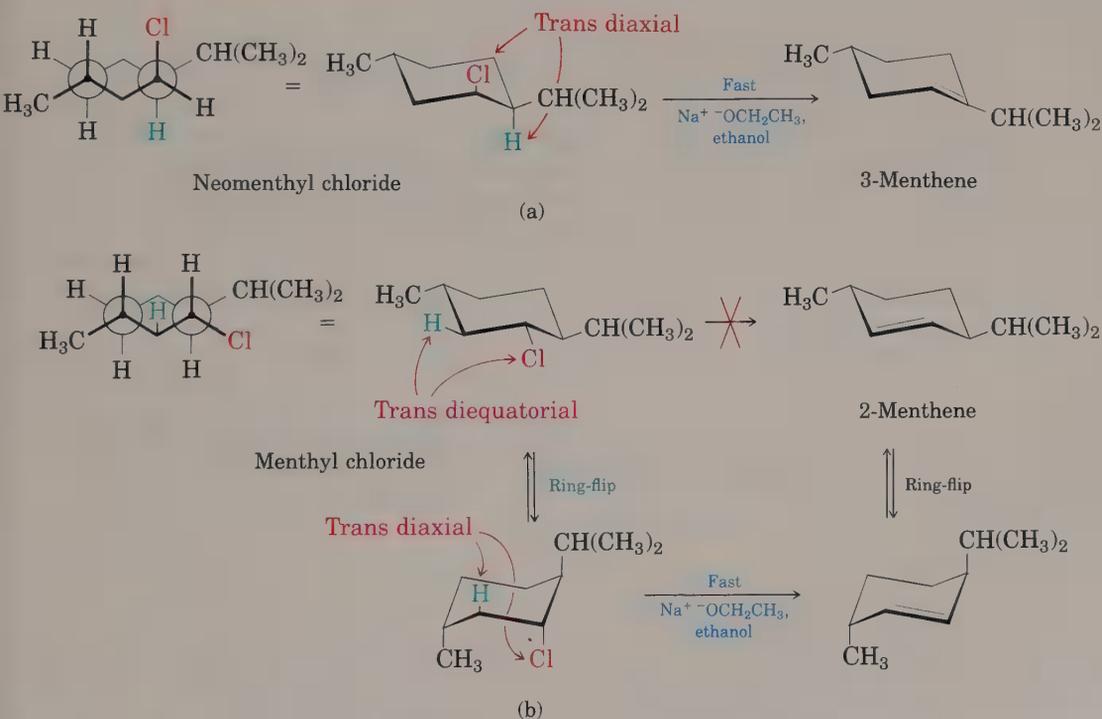
Axial chlorine: H and Cl are anti periplanar



Equatorial chlorine: H and Cl are not anti periplanar



**Figure 11.20** Geometric requirements for E2 reaction in cyclohexanes. The leaving group and the hydrogen must both be axial for anti periplanar elimination to occur.



**Figure 11.21** Dehydrochlorination of menthyl and neomenthyl chlorides. Neomenthyl chloride can lose HCl from its more stable conformation, but menthyl chloride must first ring-flip before HCl loss can occur.

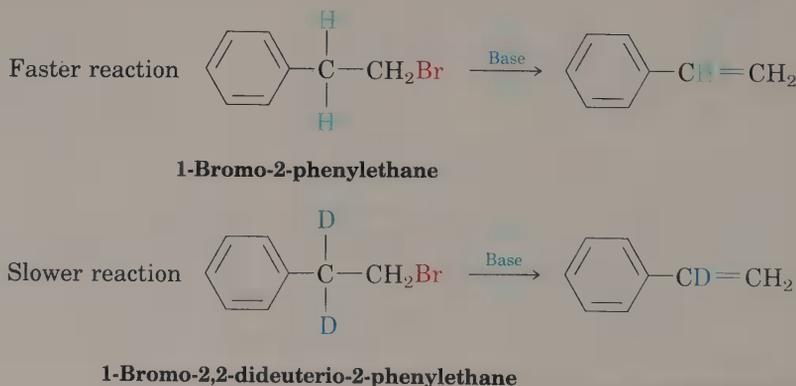
Menthyl chloride, by contrast, has the geometry shown in Figure 11.21(b), in which all three substituents are equatorial. Since the chlorine is equatorial, it can't undergo E2 elimination. In order to achieve the necessary geometry for elimination, menthyl chloride must first ring-flip to a higher-energy chair conformation, in which all three substituents are axial. Then E2 elimination occurs with loss of the only possible trans-diaxial hydrogen, leading to 2-menthene. The net effect of the simple change in chlorine stereochemistry is a 200-fold decrease in reaction rate and a complete change of product. The chemistry of the molecule is controlled by conformational effects.

PROBLEM .....

- 11.18 Which isomer would you expect to react faster under E2 elimination conditions, *trans*-1-bromo-4-*tert*-butylcyclohexane or *cis*-1-bromo-4-*tert*-butylcyclohexane? Draw each molecule in its more stable chair conformation and explain your answer.
- .....

### 11.13 The Deuterium Isotope Effect

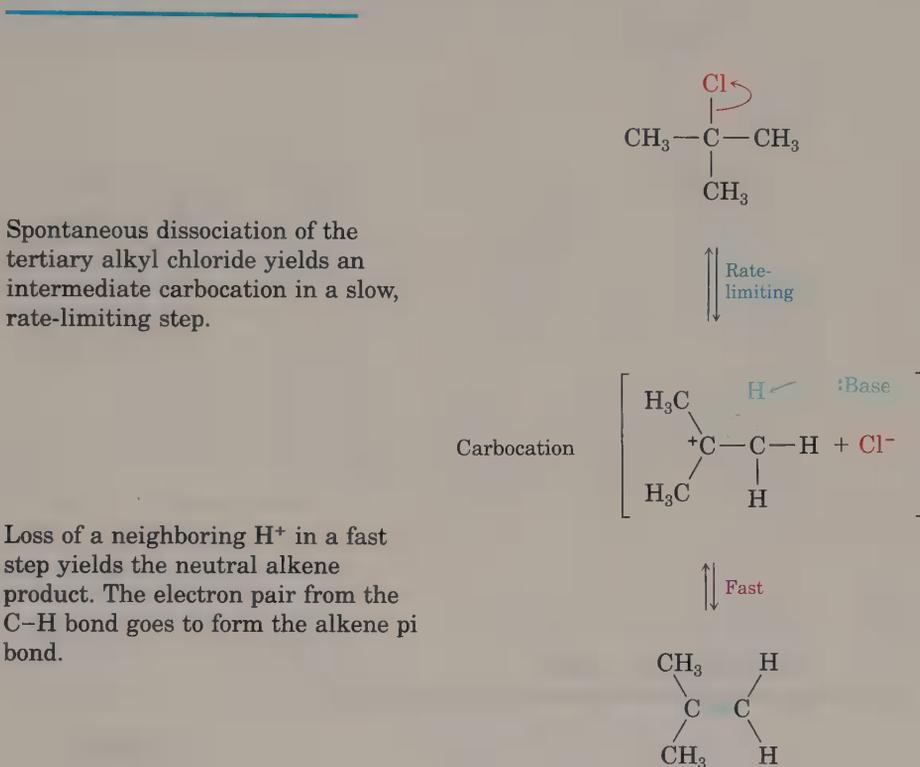
One final piece of evidence in support of the E2 mechanism is provided by a phenomenon known as the **deuterium isotope effect**. For reasons that we won't go into, a carbon–hydrogen bond is weaker by a small amount (about 1.2 kcal/mol) than a corresponding carbon–deuterium bond is. Thus, a C–H bond is more easily broken than an equivalent C–D bond, and the rate of C–H bond cleavage is therefore faster. As an example of how this effect can be used to obtain mechanistic information, the base-induced elimination of HBr from 1-bromo-2-phenylethane proceeds 7.11 times as fast as the corresponding elimination of DBr from 1-bromo-2,2-dideuterio-2-phenylethane:



This result tells us that the C–H (or C–D) bond is broken *in the rate-limiting step* and is thus fully consistent with our picture of the E2 reaction as a one-step process. If it were otherwise, we couldn't measure a rate difference.

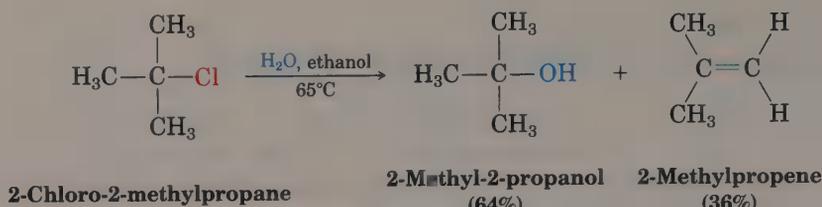
## 11.14 The E1 Reaction

Just as the  $S_N2$  reaction is analogous to the E2 reaction, the  $S_N1$  reaction also has a close analog: the **E1 reaction** (for *elimination, unimolecular*). The E1 reaction can be formulated as shown in Figure 11.22 for the elimination of HCl from 2-chloro-2-methylpropane.



**Figure 11.22** Mechanism of the E1 reaction. Two steps are involved, the first of which is rate-limiting, and a carbocation intermediate is present.

E1 eliminations begin with the same spontaneous dissociation of a halide that we saw in the  $S_N1$  reaction, but the dissociation is followed this time by loss of a proton from the intermediate carbocation. In fact, the E1 mechanism normally occurs in competition with  $S_N1$  reaction when an alkyl halide is treated in a protic solvent with a nonbasic nucleophile. Thus, the best substrates are those that are also subject to  $S_N1$  reaction, and mixtures of products are almost always obtained. For example, when 2-chloro-2-methylpropane is warmed to  $65^\circ\text{C}$  in 80% aqueous ethanol, a 64:36 mixture of 2-methyl-2-propanol ( $S_N1$ ) and 2-methylpropene (E1) is obtained:



Much evidence has been obtained in support of the E1 mechanism. As expected, E1 reactions show first-order kinetics consistent with a spontaneous dissociation process:

$$\text{Rate} = k \times [\text{RX}]$$

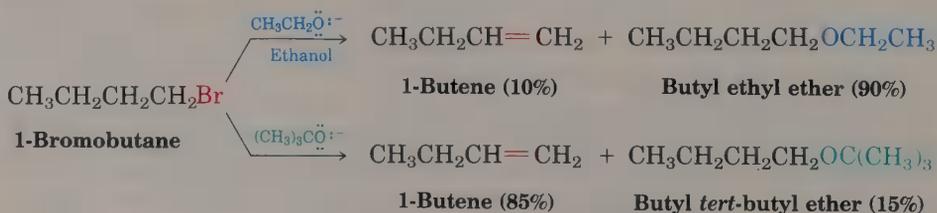
A second piece of evidence involves the stereochemistry of elimination. Unlike the E2 reaction, where periplanar geometry is required, there is no geometric requirement on the E1 reaction. The intermediate carbocation can lose any available proton from a neighboring position. We might therefore expect to obtain the more stable (Zaitsev's rule) product from E1 reaction, which is just what we find. To return to a familiar example, menthyl chloride loses HCl under E1 conditions to give a mixture of alkenes in which the Zaitsev product, 3-menthene, predominates (Figure 11.23).

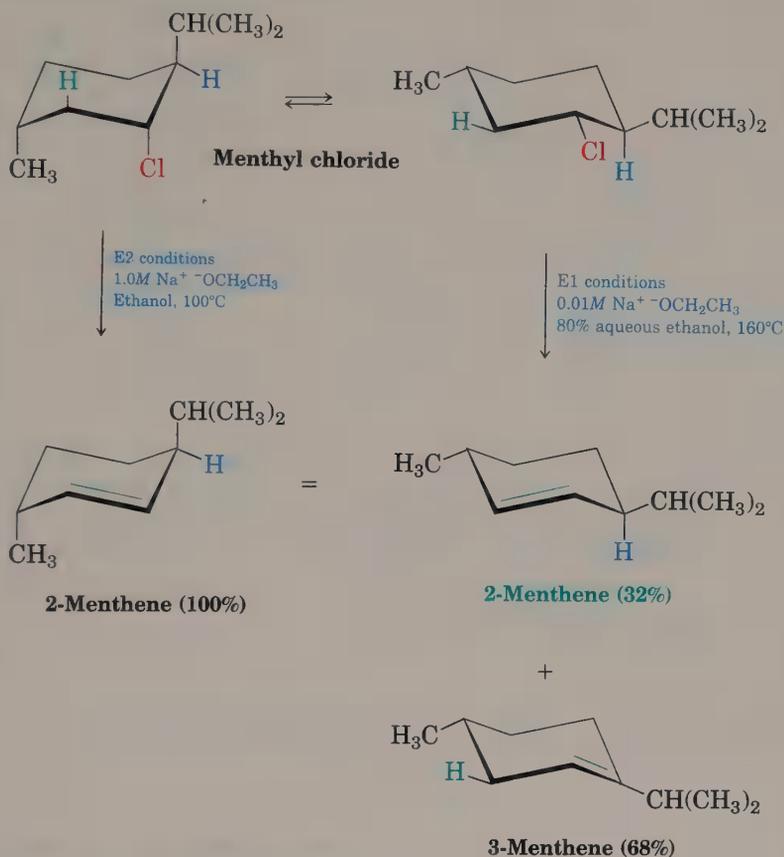
One final piece of evidence is that no deuterium isotope effect is found for E1 reactions because rupture of the C-H (or C-D) bond occurs *after* the rate-limiting step, rather than during it. Thus, we can't measure a rate difference between a deuterated and nondeuterated substrate.

## 11.15 Summary of Reactivity: $S_N1$ , $S_N2$ , E1, E2

$S_N1$ ,  $S_N2$ , E1, E2—how can you keep it all straight? How can you predict what will happen in any given case? Will substitution or elimination occur? Will the reaction be bimolecular or unimolecular? There are no rigid answers to these questions, but it *is* possible to recognize some broad trends and make some generalizations about what to expect.

1. *Primary alkyl halides* undergo substitution by an  $S_N2$  mechanism if a good nucleophile such as  $\text{RS}^-$ ,  $\text{I}^-$ ,  $\text{CN}^-$ ,  $\text{NH}_3$ , or  $\text{Br}^-$  is used. If a strong, bulky base such as *tert*-butoxide is used, E2 elimination takes place:

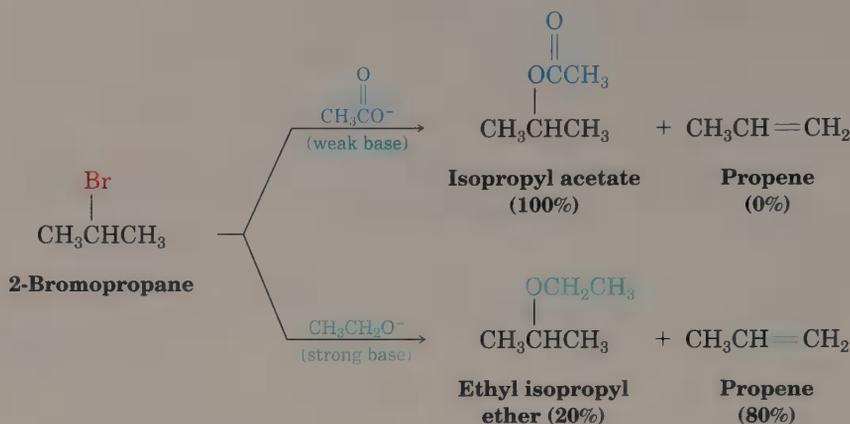




**Figure 11.23** Elimination reactions of menthyl chloride. E2 conditions (strong base) lead to 2-menthene, whereas E1 conditions (very dilute base) lead to a mixture of 2-menthene and 3-menthene.

- Secondary alkyl halides* undergo both  $S_N2$  substitution and E2 elimination, and a mixture of products often results. If a polar aprotic solvent such as hexamethylphosphoramide is used with a good nucleophile or a weak base,  $S_N2$  substitution usually predominates. If the same secondary halide is treated with a strong base such as ethoxide ion, hydroxide ion, or amide ion, however, E2 elimination occurs. For example, 2-bromopropane undergoes different reactions when treated with ethoxide ion (strong base—E2) and with acetate ion (weak base— $S_N2$ ) as seen on the next page.

Secondary alkyl halides, particularly allylic and benzylic ones, can also undergo  $S_N1$  and E1 reactions if weakly basic nucleophiles are used in protic solvents such as ethanol or acetic acid. Mixtures of products are usually obtained.



3. *Tertiary halides* undergo E2 elimination when a base such as  $\text{OH}^-$  or  $\text{RO}^-$  is used. For example, 2-bromo-2-methylpropane gives 97% elimination product when treated with ethoxide ion in ethanol. By contrast, reaction under nonbasic conditions (heating in pure ethanol) leads to a mixture of products resulting from both  $\text{S}_{\text{N}}1$  substitution and E1 elimination:

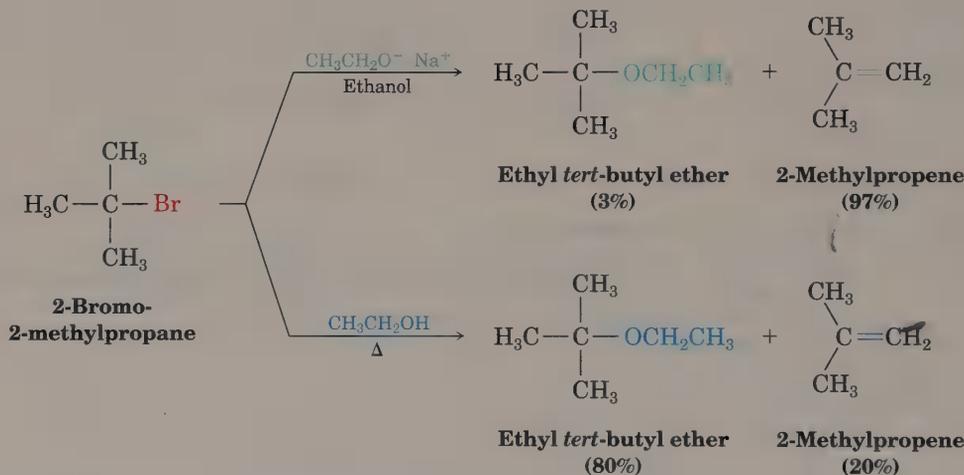
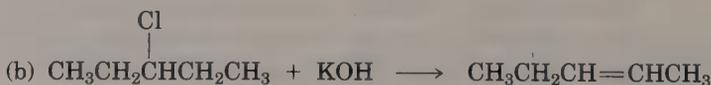
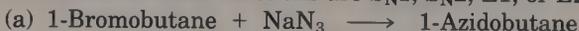
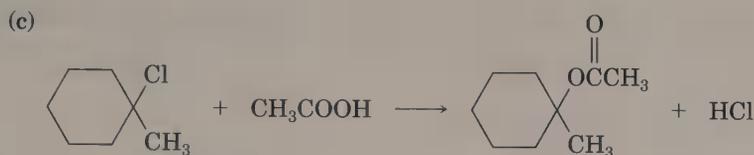


Table 11.3 summarizes these generalizations.

PROBLEM .....

11.19 Tell whether these reactions are  $\text{S}_{\text{N}}1$ ,  $\text{S}_{\text{N}}2$ , E1, or E2:





**TABLE 11.3** Correlation of Structure and Reactivity for Substitution and Elimination Reactions

Halide type	$S_N1$	$S_N2$	$E1$	$E2$
$RCH_2X$ (primary)	Does not occur	Highly favored	Does not occur	Occurs when strong bases are used
$R_2CHX$ (secondary)	Can occur with benzylic and allylic halides	Occurs in competition with $E2$ reaction	Can occur with benzylic and allylic halides	Favored when strong bases are used
$R_3CX$ (tertiary)	Favored in hydroxylic solvents	Does not occur	Occurs in competition with $S_N1$ reaction	Favored when bases are used

## 11.16 Substitution Reactions in Synthesis

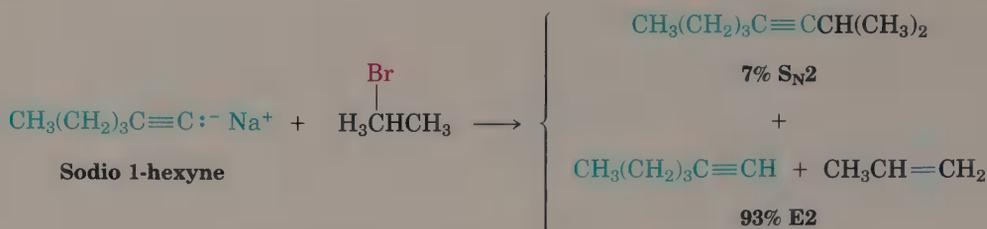
The reason we've discussed nucleophilic substitution reactions in such detail is because they're so important in organic chemistry. In fact, we've already seen a number of substitution reactions used in organic synthesis, although they weren't identified as such at the time.

One example is the alkylation of acetylide anions discussed in Section 8.9. We said that acetylide anions react well with primary alkyl bromides, iodides, and tosylates, to provide the alkyne product.



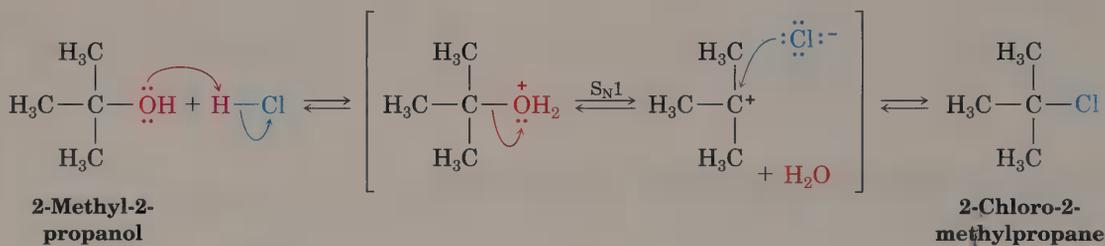
where X = Br, I, or OTos

Acetylide ion alkylation is, of course, an  $S_N2$  reaction, and it's therefore understandable that only primary alkyl halides and tosylates react well. Since acetylide anion is a strong base as well as a good nucleophile,  $E2$  elimination competes with  $S_N2$  alkylation when a secondary or tertiary substrate is used. For example, reaction of sodio 1-hexyne with 2-bromopropane gives primarily the elimination product propene:

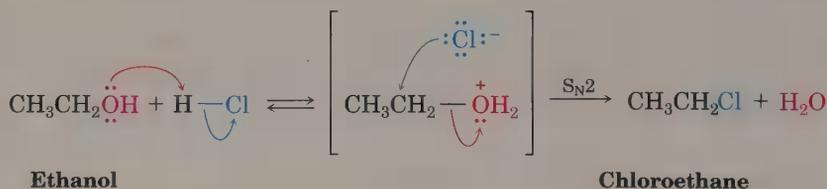


Other substitution reactions we've seen include some of the various methods used for preparing alkyl halides from alcohols. We saw in Section 10.8, for example, that alkyl halides can be prepared by treating alcohols with HX—reactions now recognizable as nucleophilic substitutions of halide ions on the protonated alcohols. Tertiary alcohols react by an S<sub>N</sub>1 pathway involving unimolecular dissociation of the protonated alcohol to yield a carbocation, whereas primary alcohols react by an S<sub>N</sub>2 pathway involving direct bimolecular displacement of halide ion on the protonated alcohol (Figure 11.24).

Tertiary alcohol—S<sub>N</sub>1



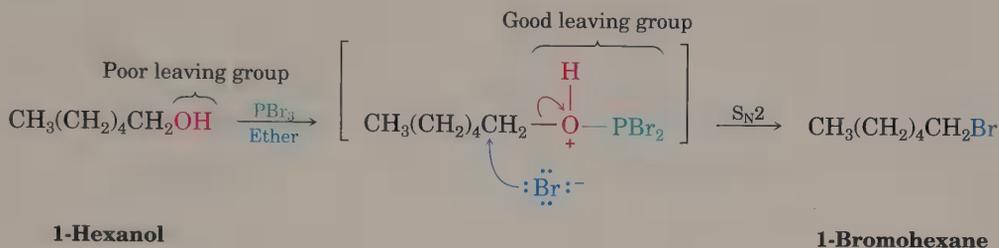
Primary alcohol—S<sub>N</sub>2



**Figure 11.24** Mechanisms of reactions of HCl with tertiary and primary alcohols.

Yet another substitution reaction we've seen is the conversion of primary and secondary alcohols into alkyl bromides by treatment with PBr<sub>3</sub> (Section 10.8). Although -OH is a poor leaving group and can't be displaced

directly by nucleophiles, reaction with  $\text{PBr}_3$  *activates* an alcohol toward displacement by transforming the hydroxyl into a better leaving group:



Alcohols react with  $\text{PBr}_3$  to give dibromophosphites ( $\text{ROPBr}_2$ ), which are highly reactive substrates in  $\text{S}_\text{N}2$  reactions. Displacement by bromide ion then occurs rapidly on the primary carbon, and alkyl bromides are produced in good yield.

## 11.17 Substitution Reactions in Biological Systems

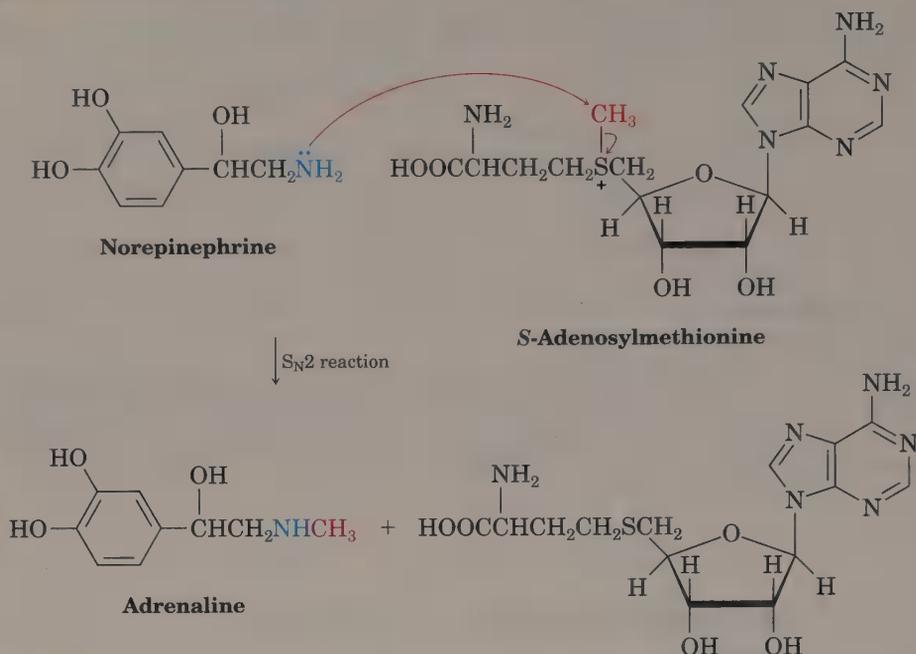
Many biological processes occur by reaction pathways analogous to those carried out in the laboratory. Thus, a number of reactions that occur in living organisms take place by nucleophilic substitution mechanisms.

### BIOLOGICAL METHYLATIONS

Perhaps the most common of all biological substitutions is the methylation reaction: the transfer of a methyl group from an electrophilic donor to a nucleophile. Although a laboratory chemist might choose iodomethane for such a reaction, living organisms operate in a more subtle way. The large and complex molecule *S*-adenosylmethionine is the biological methyl-group donor. Since the sulfur atom in *S*-adenosylmethionine has a positive charge (a *sulfonium* ion), it is an excellent leaving group for  $\text{S}_\text{N}2$  displacements on the methyl carbon. A biological nucleophile therefore attacks the methionine methyl by an  $\text{S}_\text{N}2$  reaction.

An example of the action of *S*-adenosylmethionine in biological methylations takes place in the adrenal medulla during the formation of adrenaline from norepinephrine (Figure 11.25).

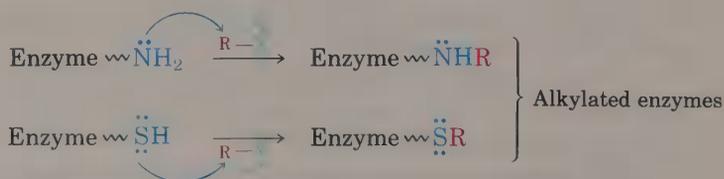
After becoming used to dealing with simple halides such as iodomethane used for laboratory alkylations, it's something of a shock to encounter a molecule as complex as *S*-adenosylmethionine. From a chemical standpoint, however,  $\text{CH}_3\text{I}$  and *S*-adenosylmethionine do exactly the same thing: Both transfer a methyl group by an  $\text{S}_\text{N}2$  reaction. The same principles of reactivity apply to both.



**Figure 11.25** The biological formation of adrenaline by  $\text{S}_\text{N}2$  reaction of norepinephrine with *S*-adenosylmethionine.

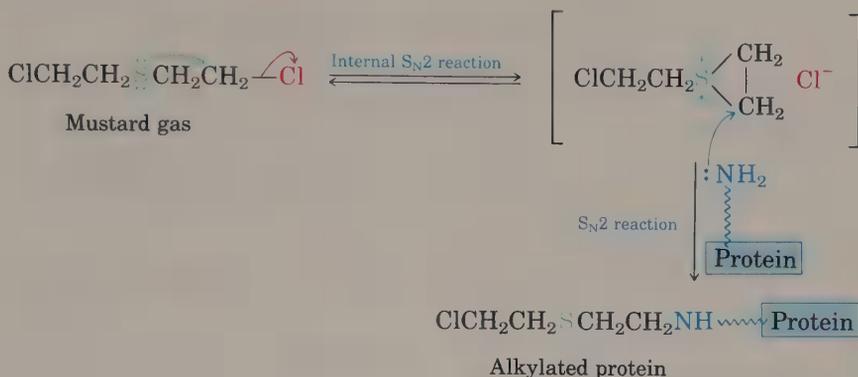
## OTHER BIOLOGICAL ALKYLATIONS

Another example of a biological  $\text{S}_\text{N}2$  reaction is involved in the response of organisms to certain toxic chemicals. Many reactive  $\text{S}_\text{N}2$  substrates with deceptively simple structures are quite toxic to living organisms. Bromomethane, for example, is widely used as a fumigant to kill termites. The toxicity of these chemicals derives from their ability to transfer alkyl groups to nucleophilic amino groups ( $-\text{NH}_2$ ) and mercapto groups ( $-\text{SH}$ ) on enzymes. With enzymes modified by alkylation, normal biological chemistry is altered.



One of the best-known toxic alkylating agents is mustard gas,  $\text{ClCH}_2\text{CH}_2\text{SCH}_2\text{CH}_2\text{Cl}$ , which gained notoriety because of its use as a chemical-warfare agent during World War I. An estimated 400,000 casualties resulted from its use. Mustard gas, a primary halide, is highly reactive toward  $\text{S}_\text{N}2$  displacements by nucleophilic groups of proteins. It is thought

to act through an intermediate sulfonium ion in much the same manner as *S*-adenosylmethionine (Figure 11.26).



**Figure 11.26** The alkylation of a protein by  $\text{S}_{\text{N}}2$  reaction of mustard gas.

## 11.18 Summary and Key Words

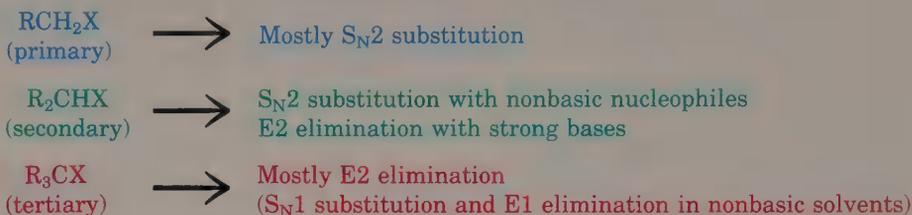
Reaction of an alkyl halide or tosylate with a nucleophile results either in **substitution** or in **elimination**. Nucleophilic substitutions are of two types:  **$\text{S}_{\text{N}}2$  reaction** and  **$\text{S}_{\text{N}}1$  reaction**. In the  $\text{S}_{\text{N}}2$  reaction, the entering nucleophile attacks the halide from a direction  $180^\circ$  away from the leaving group, resulting in an umbrella-like Walden **inversion of configuration** at the carbon atom. The reaction shows **second-order kinetics** and is strongly inhibited by increasing steric bulk of the reagents. Thus,  $\text{S}_{\text{N}}2$  reactions are favored for primary and secondary substrates.

The  $\text{S}_{\text{N}}1$  reaction occurs when the substrate spontaneously dissociates to a carbocation in a slow **rate-limiting step**, followed by a rapid attack of nucleophile. In consequence,  $\text{S}_{\text{N}}1$  reactions show **first-order kinetics** and take place with racemization of configuration at the carbon atom. They are most favored for tertiary substrates.

Eliminations of alkyl halides to yield alkenes also occur by two different mechanisms:  **$\text{E}2$  reaction** and  **$\text{E}1$  reaction**. In the  $\text{E}2$  reaction, a base abstracts a proton at the same time the leaving group departs. The reaction takes place preferentially through an **anti periplanar** transition state in which the four reacting atoms—hydrogen, two carbons, and leaving group—are all in the same plane. The reaction shows second-order kinetics and a **deuterium isotope effect**, and occurs when a secondary or tertiary substrate is treated with a strong base. These elimination reactions usually give a mixture of alkene products in which the more highly substituted alkene predominates (**Zaitsev's rule**).

The E1 reaction takes place when the substrate spontaneously dissociates to yield a carbocation in the slow rate-limiting step before losing a neighboring proton. The reaction shows first-order kinetics and no deuterium isotope effect, and occurs when a tertiary substrate reacts in polar, nonbasic solution.

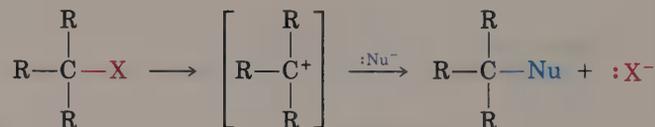
In general, substrates react in the following way:



## 11.19 Summary of Reactions

### 1. Nucleophilic substitutions

- a.  $\text{S}_{\text{N}}1$  reaction; carbocation intermediate is involved (Sections 11.6–11.9)



Best for  $3^\circ$ , allylic, and benzylic halides and tosylates.

- b.  $\text{S}_{\text{N}}2$  reaction; back-side attack of nucleophile occurs (Sections 11.4–11.5)

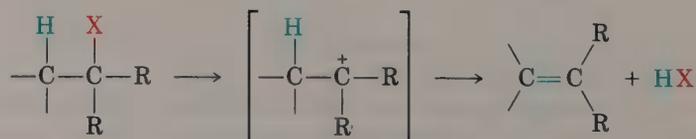


Best for  $1^\circ$  or  $2^\circ$  halides

$\text{Nu}:\text{---} = \text{H}^-, \text{CN}^-, \text{I}^-, \text{Br}^-, \text{Cl}^-, \text{OH}^-, \text{NH}_2^-, \text{CH}_3\text{O}^-, \text{CH}_3\text{CO}_2^-, \text{HS}^-, \text{H}_2\text{O}, \text{NH}_3$ , and so forth.

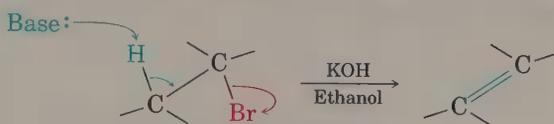
### 2. Eliminations

- a. E1 reaction; more highly substituted alkene is formed (Section 11.14)



Best for  $3^\circ$  halides

b. E2 reaction; anti periplanar geometry is required (Section 11.11)



Best for 2° and 3° halides

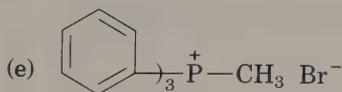
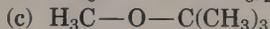
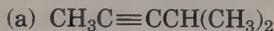
## ADDITIONAL PROBLEMS

- 11.20 Which reagent in each pair will react faster in an  $S_N2$  reaction with hydroxide ion?
- $\text{CH}_3\text{Br}$  or  $\text{CH}_3\text{I}$
  - $\text{CH}_3\text{CH}_2\text{I}$  in ethanol or dimethyl sulfoxide
  - $(\text{CH}_3)_3\text{CCl}$  or  $\text{CH}_3\text{Cl}$
  - $\text{H}_2\text{C}=\text{CHBr}$  or  $\text{H}_2\text{C}=\text{CHCH}_2\text{Br}$

- 11.21 What effect would you expect these changes to have on the rate of the reaction of 1-iodo-2-methylbutane with cyanide ion?
- $\text{CN}^-$  concentration is halved, and 1-iodo-2-methylbutane concentration is doubled.
  - Both  $\text{CN}^-$  and 1-iodo-2-methylbutane concentrations are tripled.

- 11.22 What effect would you expect on the rate of reaction of ethanol with 2-iodo-2-methylbutane if the concentration of the halide were tripled?

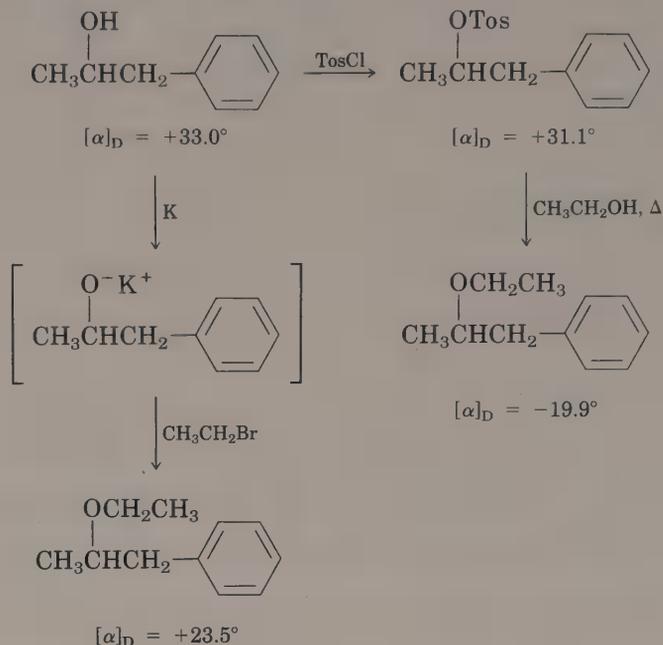
- 11.23 How might you prepare each of the following molecules using a nucleophilic substitution reaction at some step?



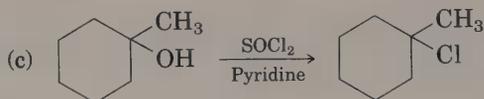
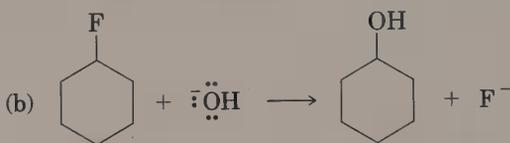
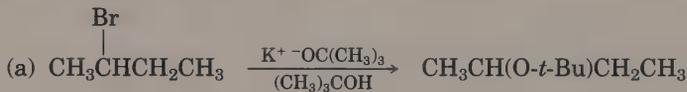
A phosphonium salt

- 11.24 Which reaction in each of these pairs would you expect to be faster?
- The  $S_N2$  displacement by iodide ion on  $\text{CH}_3\text{Cl}$  or on  $\text{CH}_3\text{OTos}$
  - The  $S_N2$  displacement by acetate ion on bromoethane or on bromocyclohexane
  - The  $S_N2$  displacement on 2-bromopropane by ethoxide ion or by cyanide ion
  - The  $S_N2$  displacement by acetylide ion on bromomethane in benzene or in hexamethylphosphoramide
- 11.25 What products would you expect from the reaction of 1-bromopropane with:
- $\text{NaNH}_2$
  - $\text{K}^+ \text{ } ^-\text{OC}(\text{CH}_3)_3$
  - $\text{NaI}$
  - $\text{NaCN}$
  - $\text{Na}^+ \text{ } ^-\text{C}\equiv\text{CH}$
  - $\text{Mg}$ , then  $\text{H}_2\text{O}$
- 11.26 Which reagent in each of the following pairs is more nucleophilic? Explain your answer.
- $\text{ } ^-\text{NH}_2$  or  $\text{NH}_3$
  - $\text{H}_2\text{O}$  or  $\text{CH}_3\text{COO}^-$
  - $\text{BF}_3$  or  $\text{F}^-$
  - $(\text{CH}_3)_3\text{P}$  or  $(\text{CH}_3)_3\text{N}$
  - $\text{I}^-$  or  $\text{Cl}^-$
  - $\text{ } ^-\text{C}\equiv\text{N}$  or  $\text{ } ^-\text{OCH}_3$

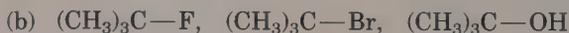
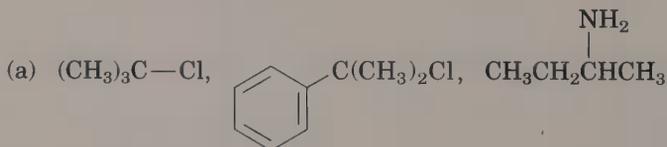
- 11.27 Among the Walden cycles carried out by Phillips and Kenyon is the following series of reactions reported in 1923. Explain the results and indicate where Walden inversion is occurring.

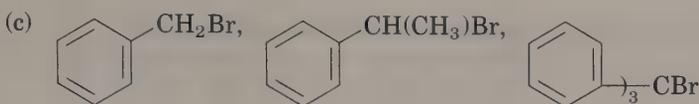


- 11.28 The synthetic sequences shown here are unlikely to occur as written. What is wrong with each?

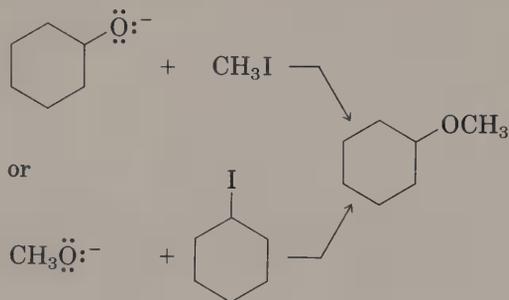


- 11.29 Order each set of compounds with respect to  $\text{S}_{\text{N}}1$  reactivity.

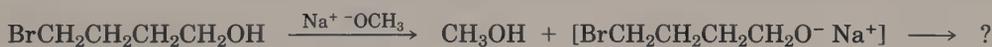




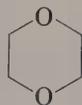
- 11.30** Order each set of compounds with respect to  $S_N2$  reactivity.
- (a)  $(CH_3)_3CCl$ ,  $CH_3CH_2CH_2Cl$ ,  $CH_3CH_2CHClCH_3$   
 (b)  $(CH_3)_2CHCHBrCH_3$ ,  $(CH_3)_2CHCH_2Br$ ,  $(CH_3)_3CCH_2Br$   
 (c)  $CH_3CH_2CH_2OCH_3$ ,  $CH_3CH_2CH_2OTos$ ,  $CH_3CH_2CH_2Br$
- 11.31** Predict the product and give the stereochemistry resulting from reaction of the following nucleophiles with (*R*)-2-bromooctane:
- (a)  $^-CN$                       (b)  $CH_3CO_2^-$                       (c)  $CH_3S^-$                       (d)  $Br^-$
- 11.32** Describe the effects of the following variables on  $S_N2$  and  $S_N1$  reactions:
- (a) Solvent    (b) Leaving group  
 (c) Attacking nucleophile    (d) Substrate structure
- 11.33** Ethers can often be prepared by  $S_N2$  reaction of alkoxide ions with alkyl halides. Suppose you wanted to prepare cyclohexyl methyl ether. Which of the two possible routes shown here would you choose? Explain.



- 11.34** The  $S_N2$  reaction can occur *intramolecularly* (within the same molecule). What product would you expect from treatment of 4-bromo-1-butanol with base?

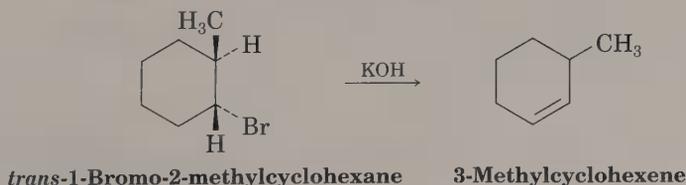


- 11.35** In light of your answer to Problem 11.34, propose a synthesis of 1,4-dioxane starting only with a dihalide.

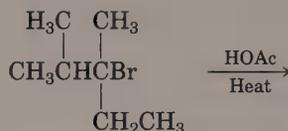


1,4-Dioxane

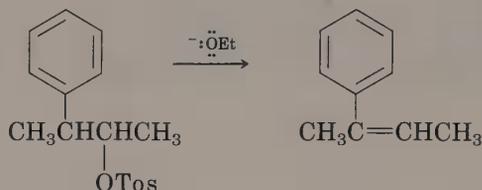
- 11.36** How can you explain the fact that *trans*-1-bromo-2-methylcyclohexane yields the non-Zaitsev elimination product 3-methylcyclohexene on treatment with base?



- 11.37 Propose structures for compounds that fit these descriptions:  
 (a) An alkyl halide that gives a mixture of three alkenes on E2 reaction  
 (b) An alkyl halide that will not undergo nucleophilic substitution  
 (c) An alkyl halide that gives the non-Zaitsev product on E2 reaction  
 (d) An alcohol that reacts rapidly with HCl at 0°C
- 11.38 Predict the major alkene product from this elimination:

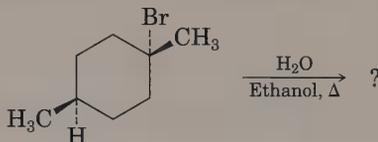


- 11.39 The tosylate of (2*R*,3*S*)-3-phenyl-2-butanol undergoes E2 elimination on treatment with ethoxide ion to yield (*Z*)-2-phenyl-2-butene.

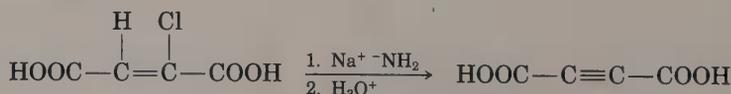


Formulate the reaction, showing the proper stereochemistry. Explain the observed result by using Newman projections.

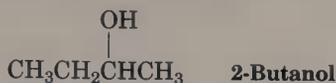
- 11.40 In light of your answer to Problem 11.39, which alkene, *E* or *Z*, would you expect from the elimination reaction of the (2*R*,3*R*)-3-phenyl-2-butanol tosylate? Which alkene would result from E2 reaction on the 2*S*,3*R* and 2*S*,3*S* tosylates? Explain your answers.
- 11.41 Predict the product(s) of the following reaction, making sure you indicate stereochemistry where necessary.



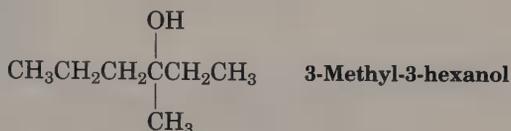
- 11.42 Draw all isomers of  $\text{C}_4\text{H}_9\text{Br}$ , name them, and arrange them in order of decreasing reactivity in the  $\text{S}_{\text{N}}2$  reaction.
- 11.43 Reaction of iodoethane with cyanide ion,  $\text{CN}^-$ , yields primarily the nitrile  $\text{CH}_3\text{CH}_2\text{C}\equiv\text{N}$  along with a small amount of isonitrile,  $\text{CH}_3\text{CH}_2\text{N}\equiv\text{C}$ . Write Lewis structures for both products, assign formal charges as necessary, and propose mechanisms to account for their formation.
- 11.44 Alkynes can be made by dehydrohalogenation of vinylic halides in a reaction that is essentially an E2 process. In studying the stereochemistry of this elimination, it was found that (*Z*)-2-chloro-2-butene-1,4-dioic acid reacts 50 times as fast as the corresponding *E* isomer. What conclusion can you draw about the stereochemistry of eliminations in vinylic halides? How does this result compare with eliminations of alkyl halides?



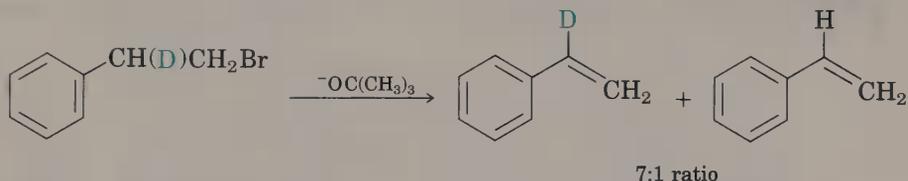
- 11.45 Optically active 2-butanol slowly racemizes on standing in dilute sulfuric acid. Propose a mechanism to account for this observation.



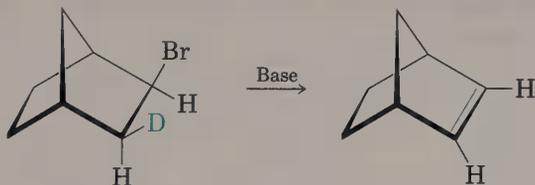
- 11.46 Suggest an explanation for the observation that reaction of HBr with (*R*)-3-methyl-3-hexanol leads to ( $\pm$ )-3-bromo-3-methylhexane.



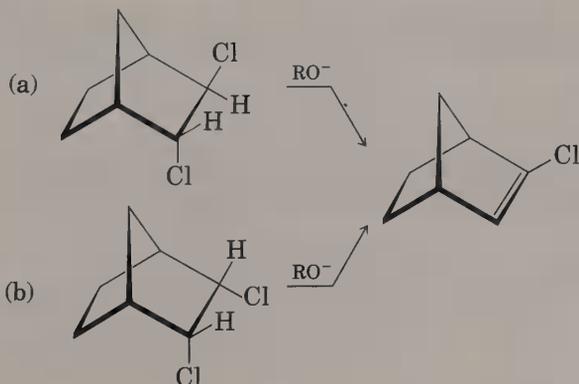
- 11.47 Explain the fact that treatment of 1-bromo-2-deuterio-2-phenylethane with strong base leads to a mixture of deuterated and nondeuterated phenylethylenes in which the deuterated product predominates by approximately 7:1.



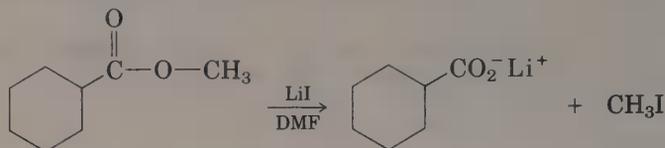
- 11.48 Although anti geometry is preferred for E2 reactions, it isn't absolutely necessary. The deuterated bromo compound shown here reacts with strong base to yield an undeuterated alkene. Clearly, a syn elimination has occurred. Make a molecular model of the starting material to examine its geometry, and then explain the result.



- 11.49 In light of your answer to Problem 11.48, account for the observation that one of the following isomers undergoes E2 reaction approximately 100 times as fast as the other. Which isomer is the more reactive, and why?

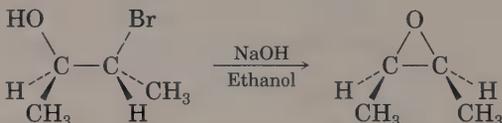


- 11.50 There are eight possible diastereomers of 1,2,3,4,5,6-hexachlorocyclohexane. Draw them in their most stable chair conformations. One isomer loses HCl in an E2 reaction nearly 1000 times more slowly than the others. Which isomer reacts so slowly, and why?
- 11.51 Consider the following methyl ester cleavage reaction:

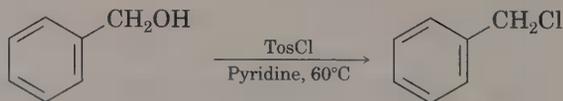


The following evidence is available:

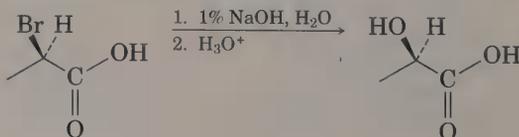
- (a) The reaction occurs much faster in DMF than in ethanol.  
 (b) The corresponding ethyl ester cleaves approximately 10 times more slowly than the methyl ester.
- Using this evidence, propose a mechanism for the reaction. What other kinds of experimental evidence could you gather to support your mechanistic hypothesis?
- 11.52 Account for the fact that the rate of reaction of 1-chlorooctane with acetate ion to give octyl acetate is greatly accelerated by the presence of a small quantity of iodide ion.
- 11.53 Compound X is optically inactive and has the formula  $\text{C}_{16}\text{H}_{16}\text{Br}_2$ . On treatment with strong base, X gives hydrocarbon Y,  $\text{C}_{16}\text{H}_{14}$ . Compound Y absorbs two equivalents of hydrogen when reduced over a palladium catalyst and reacts with ozone to give two fragments. One fragment, Z, is an aldehyde with formula  $\text{C}_7\text{H}_6\text{O}$ . The other fragment is glyoxal,  $\text{CHOCHO}$ . Formulate the reactions involved, and suggest structures for X, Y, and Z. What is the stereochemistry of X?
- 11.54 Propose a structure for an alkyl halide that gives (*E*)-3-methyl-2-phenyl-2-pentene and none of the *Z* isomer on E2 elimination. Make sure you indicate the stereochemistry.
- 11.55 The following reaction is simply a nucleophilic substitution, even though you haven't seen it before. Propose a mechanism to account for the observed result.



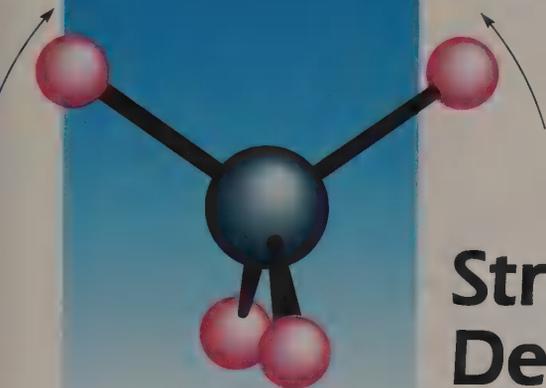
- 11.56 The following reaction is simply a nucleophilic substitution, even though you haven't seen it before. Propose a mechanism to account for the observed result.



- 11.57 The following reaction is simply a nucleophilic substitution, even though you haven't seen it before. Propose a mechanism to account for the observed result.



# Chapter 12



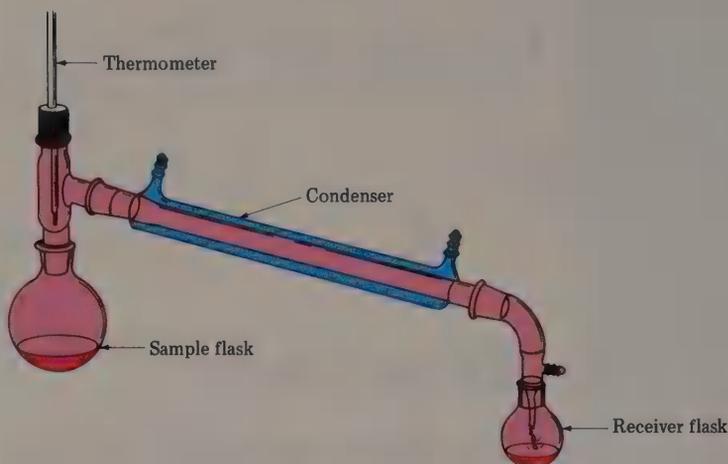
## Structure Determination: Mass Spectroscopy and Infrared Spectroscopy

Structure determination is central to organic chemistry. Every time a reaction is run, the products must be purified and identified. The products must be separated from solvents, from excess reagents, and from each other if more than one is formed, and their structures must be determined. Determining the structure of an organic compound was a time-consuming process in the nineteenth and early twentieth centuries, but extraordinary advances have been made in the past few decades. Sophisticated instruments are now available that greatly simplify the problems of purification and structure determination. Use of these instruments doesn't guarantee good results—skill and patience are still required—but it does make things easier.

### 12.1 Purification of Organic Compounds

**Crystallization** is one of the simplest yet most effective methods for purifying a solid product. The crude reaction product is dissolved in a minimum amount of a suitable hot solvent, and the solution is allowed to cool slowly. Pure crystals slowly form and precipitate, while impurities remain in solution. The crystalline product is then isolated by filtration.

**Distillation** is an equally simple and effective method for purifying a volatile liquid product. The crude liquid reaction product is heated to a boil, and the vapors are condensed into a receiver flask (Figure 12.1). Nonvolatile impurities remain in the still pot.



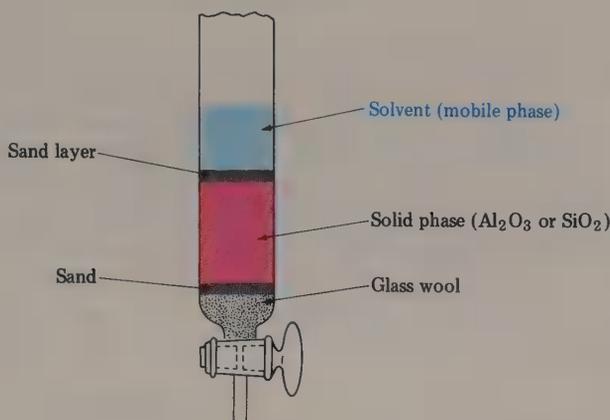
**Figure 12.1** A simple distillation apparatus. As the volatile sample is boiled, vapors rise, are condensed, and are collected in a receiver flask.

If the crude product is a mixture of two or more volatile compounds having different boiling points, **fractional distillation** can often effect a separation. The lower-boiling, more volatile component distills first, followed by the higher-boiling material. The theory of fractional distillation columns is complex, but as a practical matter good separation can usually be achieved in the laboratory if the components differ in boiling point by more than  $10^{\circ}\text{C}$ .

If neither crystallization nor distillation is effective, some form of **chromatography** is normally used to separate a mixture of organic compounds. The development of chromatography, which literally means “color writing,” as a separation technique dates back to the work of the Russian chemist Mikhail Tswett<sup>1</sup> in 1903. Tswett described the separation of the pigments in green leaves by dissolving the leaf extract in an organic solvent and allowing the solution to run down through a vertical glass tube packed with chalk powder. Different pigments passed down the column at different rates, leaving a series of colored bands on the white chalk column. A simple chromatographic column is shown in Figure 12.2.

The term *chromatography* is now used to refer to a variety of related separation techniques, all of which work on a common principle: The mixture to be separated is dissolved in a solvent **mobile phase** and passed over an adsorbent **stationary phase**. Because different compounds adsorb to the stationary phase to differing extents, they migrate through the phase at different rates and are separated as they emerge (**elute**) from the end of the

<sup>1</sup>Mikhail Semenovich Tswett (or Tsvett) (1872–1919); b. Asti, Italy; Russian subject; University of Warsaw (1901–1907); Institute of Technology, Warsaw (1908); in Moscow (1915); professor and director, Botanical Garden, Estonia (1917).



**Figure 12.2** A simple chromatography column. The glass tube is filled with adsorbent material, and a solution containing the mixture to be separated is allowed to drip through the column.

chromatography column. We'll discuss briefly three chromatographic techniques that are often used by organic chemists: liquid chromatography, high-performance liquid chromatography, and gas chromatography.

## 12.2 Liquid Chromatography

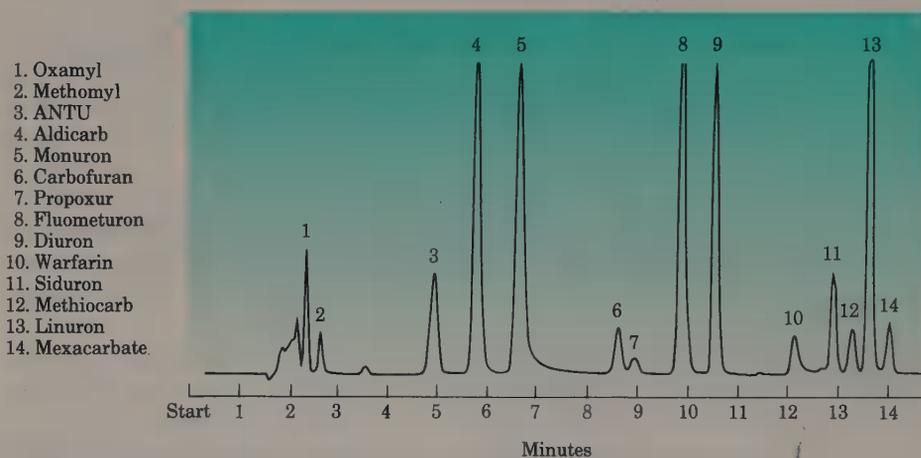
**Liquid chromatography**, or **column chromatography**, is one of the simplest and most often used chromatographic methods. As in Tswett's original experiments, a mixture of organic compounds is dissolved in a suitable solvent (the mobile phase) and is adsorbed onto a stationary phase such as alumina ( $\text{Al}_2\text{O}_3$ ) or silica gel (hydrated  $\text{SiO}_2$ ). More solvent is then passed down the column containing the stationary phase, and different compounds are eluted at different times.

The time at which a compound is eluted is strongly influenced by its polarity. As a general rule, molecules with polar functional groups are adsorbed more strongly and therefore migrate through the stationary phase more slowly than nonpolar molecules. For example, a mixture of an alcohol such as 1-heptanol and a related alkene such as 1-heptene can be easily separated by liquid chromatography. The relatively nonpolar alkene passes through the column much faster than the more polar alcohol.

**High-performance liquid chromatography (HPLC)** is a variant of the simple column technique. It has been found that the efficiency of column chromatography is vastly improved if the stationary phase is made up of very small, uniformly sized spherical particles. Small particle size ensures a large surface area for better adsorption, and a uniform spherical shape allows a tight, uniform packing. In practice, specially prepared and coated

silica microspheres of 10–25 micrometer ( $\mu\text{m}$ ) size ( $1 \mu\text{m} = 10^{-6} \text{ m}$ ) are often used: 15 g of these microspheres have a surface area equivalent to the size of a football field.

High-pressure pumps are required to force solvent through a tightly packed HPLC column, and sophisticated detectors are required to monitor the appearance of material eluting from the column. These refinements are well worth it, however, for a good HPLC column can have up to several thousand times the separating power of a simple gravity-flow column. Figure 12.3 shows the results of HPLC analysis of a mixture of 14 common pesticides, using coated silica microspheres as the stationary phase and acetonitrile/water as the mobile phase. As each compound is eluted from the chromatography column, it passes through a detector, which registers its presence as a peak on a recorder chart.

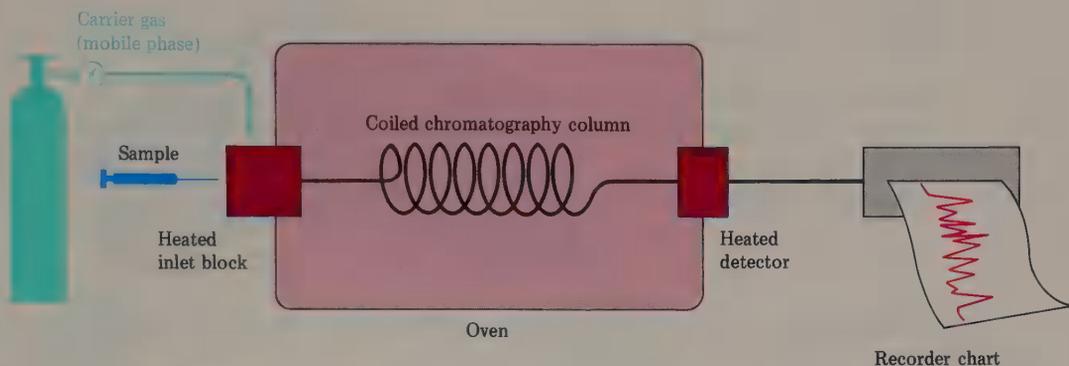


**Figure 12.3** The HPLC analysis of a mixture of 14 agricultural pesticides. The structures of the pesticides can be found in the *Merck Index*.

## 12.3 Gas Chromatography

**Gas chromatography** differs from liquid chromatography in several respects. The most important difference is that a *gas* such as nitrogen or helium is used as the mobile phase, rather than a liquid solvent. Operationally, the technique employs an instrument known as a gas chromatograph, shown schematically in Figure 12.4.

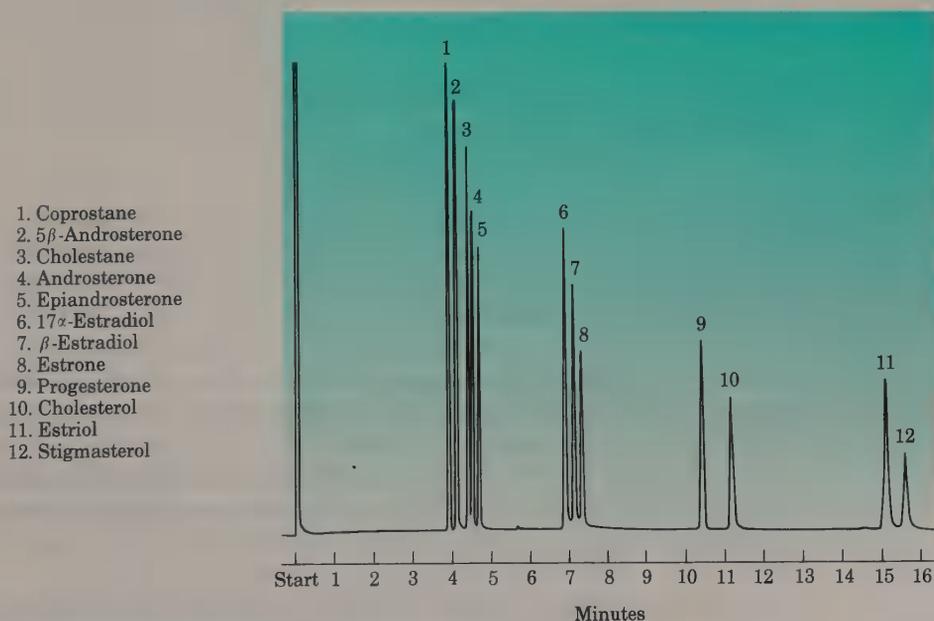
A small amount of sample mixture (often less than 1 mg) is dissolved in a small volume of solvent and injected by syringe into a heated inlet of the gas chromatograph. The sample is instantly vaporized and is then swept through a heated chromatography column by a stream of carrier gas. As



**Figure 12.4** Schematic representation of a gas chromatograph.

pure separated components elute from the column, the appearance of each is detected and registered as a peak on a recorder chart.

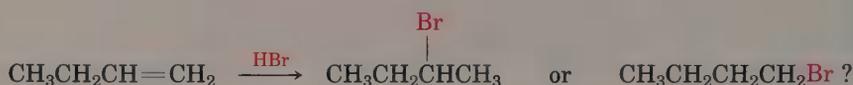
Although only small amounts of material can be separated on a gas chromatograph, the separating power of modern instruments is phenomenal. Efficiencies up to 10 times that of HPLC and 100,000 times that of simple laboratory distillation can be achieved. Figure 12.5 shows the results of an analysis carried out on a mixture of 12 steroids, using a 30 m fused-silica (quartz) column.



**Figure 12.5** Gas chromatographic analysis of a steroid mixture. The structures can be found in the *Merck Index*.

## 12.4 Structure Determination: Spectroscopy

Once the products of a reaction have been separated and purified, the more difficult work begins. How do we identify what's been made? How do we know, for example, that the ionic addition of HBr to 1-butene gives 2-bromobutane and not 1-bromobutane?



The answer is that we use several different kinds of *spectroscopy* to elucidate the structures of unknowns. In this and the next two chapters we'll look at four of the most useful spectroscopic techniques—mass spectroscopy (MS), infrared spectroscopy (IR), nuclear magnetic resonance spectroscopy (NMR), and ultraviolet spectroscopy (UV)—and we'll see the kind of information to be gotten from each.

**Mass spectroscopy**

What size and formula?

**Infrared spectroscopy**

What functional groups are present?

**Ultraviolet spectroscopy**

Is a conjugated pi electron system present?

**Nuclear magnetic resonance spectroscopy**

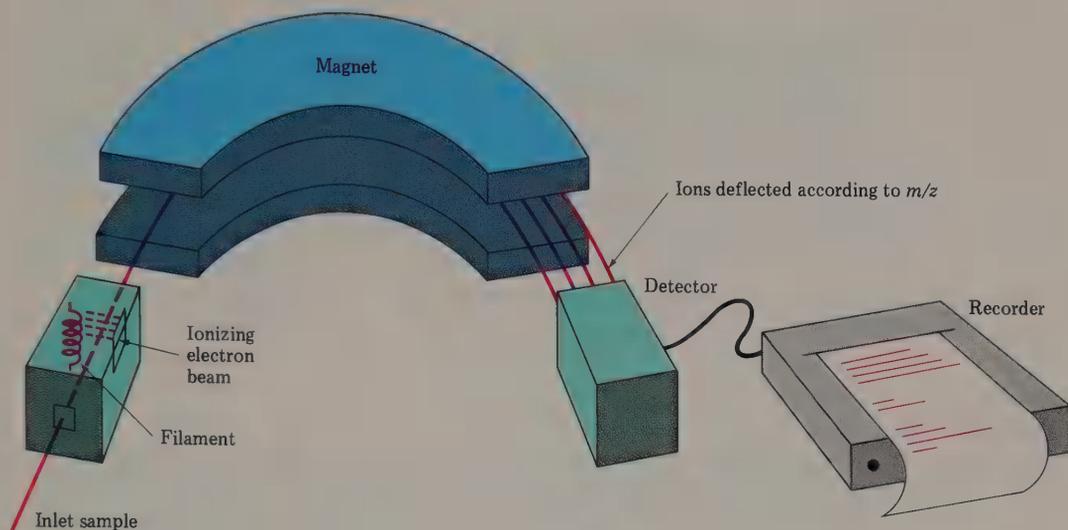
What carbon–hydrogen framework is present?

Normally, no one single spectroscopic method is reliable by itself for determining the structure of an unknown. A combination of two or more methods is almost always used.

## 12.5 Mass Spectroscopy

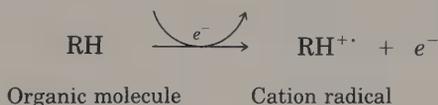
At its simplest, **mass spectroscopy (MS)** is a technique that allows us to measure the mass (molecular weight) of a molecule. In addition, we can often gain valuable structural information about unknowns by measuring the masses of the fragments produced when high-energy molecules fly apart. There are several different kinds of mass spectrometers available, but one of the most common is the electron-impact, magnetic-sector instrument shown schematically in Figure 12.6.

A small amount of sample is introduced into the mass spectrometer, where it is bombarded by a stream of high-energy electrons. The exact energy of the electron beam can be varied but is commonly around 70 electron volts (eV), or 1600 kcal/mol (6700 kJ/mol). When a high-energy electron



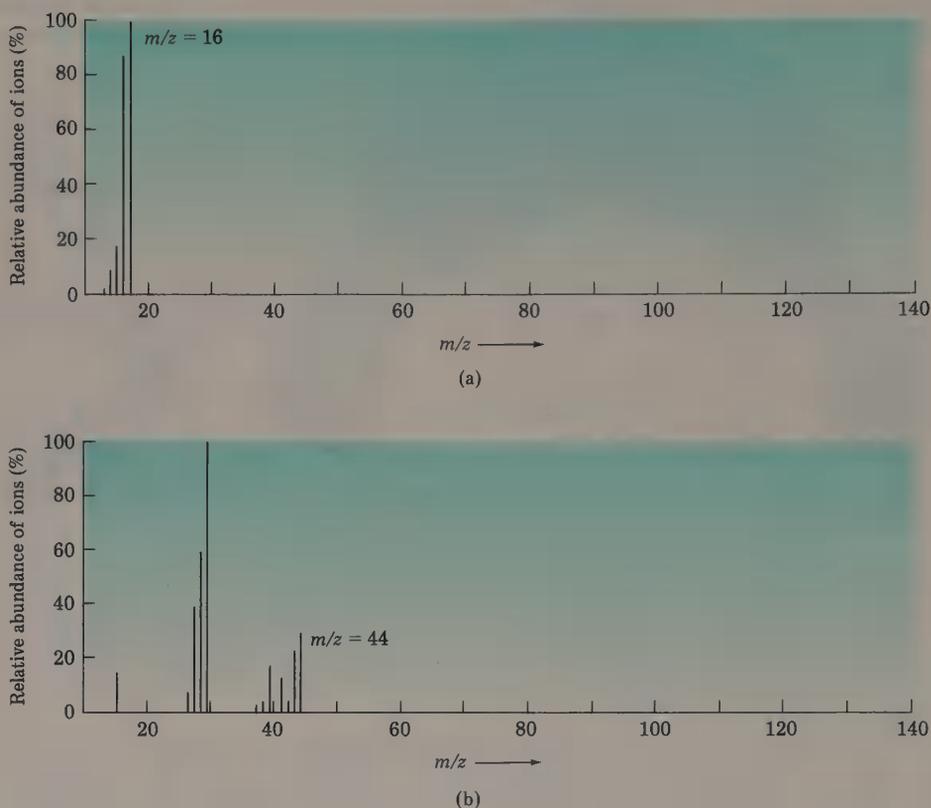
**Figure 12.6** Schematic representation of an electron-impact, magnetic-sector mass spectrometer.

strikes an organic molecule, it dislodges a valence electron from the molecule, producing a cation radical—*cation* because the molecule has lost a negatively charged electron; *radical* because the molecule now has an odd number of electrons.



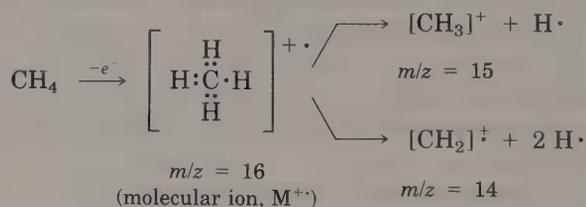
Electron bombardment transfers such a large amount of energy to the sample molecules that the cation radicals *fragment* after ionization: They fly apart into numerous smaller pieces, some of which retain a positive charge, and some of which are neutral. The fragments then pass through a strong magnetic field, where they are deflected through a curved pipe according to their mass-to-charge ratio ( $m/z$ ). Neutral fragments are not deflected by the magnetic field and are lost on the walls of the pipe, but positively charged fragments are sorted by the mass spectrometer onto a detector, which records them as peaks at the proper  $m/z$  ratios. Since the number of charges,  $z$ , is usually 1, the peaks of ratio  $m/z$  are simply  $m$ , the masses of the ions.

The **mass spectrum** of a compound is usually presented as a bar graph with unit masses ( $m/z$  values) on the  $x$ -axis and intensity (number of ions of a given  $m/z$  striking the detector) on the  $y$ -axis. The tallest peak, called the **base peak**, is arbitrarily assigned an intensity of 100%. Figure 12.7 shows mass spectra of methane and propane.



**Figure 12.7** Mass spectra of (a) methane (CH<sub>4</sub>; mol wt = 16) and (b) propane (C<sub>3</sub>H<sub>8</sub>; mol wt = 44).

The mass spectrum of methane is relatively simple, since few fragmentations are possible. As Figure 12.7(a) shows, the base peak has  $m/z = 16$ , which corresponds to the unfragmented methane cation radical, called the **parent peak**, or the **molecular ion (M<sup>+</sup>)**. The mass spectrum also shows ions at  $m/z = 15$  and  $14$ , corresponding to cleavage of the molecular ion into CH<sub>3</sub><sup>+</sup> and CH<sub>2</sub><sup>+</sup> fragments.



The mass spectral fragmentation patterns of larger molecules are usually complex, and the molecular ion is often not the highest (base) peak. For example, the mass spectrum of propane, shown in Figure 12.7(b), has

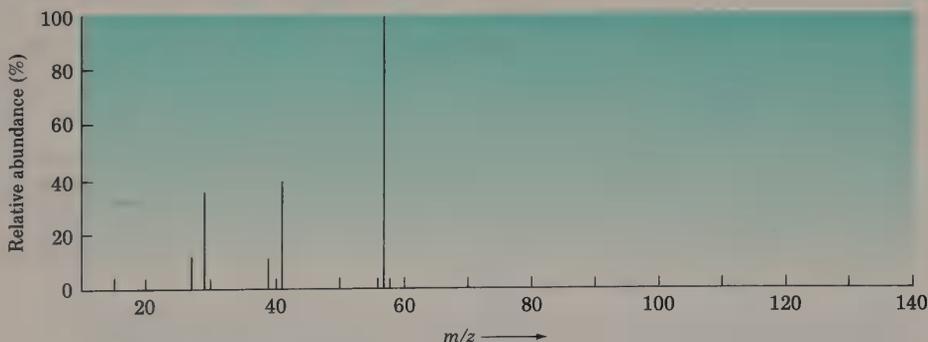
a molecular ion at  $m/z = 44$  that is only about 30% as high as the base peak at  $m/z = 29$ . In addition, many other fragment ions are observed.

## 12.6 Interpreting Mass Spectra

What kinds of information can be gained by studying the mass spectrum of a compound? Certainly the most obvious piece of information is the molecular weight (mol wt), which in itself can be invaluable. For example, if we were given three unlabeled bottles containing hexane (mol wt = 86), 1-hexene (mol wt = 84), and 1-hexyne (mol wt = 82), mass spectroscopy would readily distinguish among them.

Some instruments have such sophisticated detectors that they provide mass measurements accurate to 0.001 atomic mass unit (amu), making it possible to distinguish between two formulas with the same nominal mass. For example, although both  $C_5H_{12}$  and  $C_4H_8O$  have mol wt = 72, they differ slightly beyond the decimal point:  $C_5H_{12}$  has an exact mass of 72.0939, whereas  $C_4H_8O$  has an exact mass of 72.0575. A high-resolution instrument would easily distinguish between the two.

Unfortunately, not all compounds show a molecular ion in their mass spectra. Although  $M^+$  is usually easy to identify if it's abundant, some compounds such as 2,2-dimethylpropane fragment so readily that no molecular ion is observed (Figure 12.8).



**Figure 12.8** Mass spectrum of 2,2-dimethylpropane ( $C_5H_{12}$ ; mol wt = 72). No molecular ion is observed.

Knowing an unknown's molecular weight makes it possible to narrow the choices for molecular formula down to only a few. For example, if the mass spectrum of an unknown compound shows a molecular ion at  $m/z = 110$ , the molecular formula is likely to be  $C_8H_{14}$ ,  $C_7H_{10}O$ ,  $C_6H_6O_2$ , or  $C_6H_{10}N_2$ . There are always a number of possible molecular formulas for all but the lowest molecular weights, and tables are available that list all possible choices.

A further point about mass spectroscopy can be seen by looking carefully at the mass spectra of methane and propane in Figure 12.7. Surprisingly, the molecular ions aren't the highest mass peaks in the two spectra; there are small peaks in each spectrum at  $(M + 1)^+$  due to the presence in the samples of small amounts of isotopically substituted molecules. Although  $^{12}\text{C}$  is the most abundant carbon isotope, a small amount (1.11% natural abundance) of  $^{13}\text{C}$  is also present. Thus, a certain percentage of the molecules analyzed in the mass spectrometer are likely to contain a  $^{13}\text{C}$  isotope, giving rise to the observed  $(M + 1)^+$  peak.

PROBLEM .....

12.1 Write as many molecular formulas as you can for compounds that have the following molecular ions in their mass spectra. Assume that all of the compounds contain C and H, but that O may or may not be present.

(a)  $M^{+} = 86$

(b)  $M^{+} = 128$

(c)  $M^{+} = 156$

PROBLEM .....

12.2 Nootkatone, one of the chemicals responsible for the odor and taste of grapefruit, shows a molecular ion at  $m/z = 218$  in its mass spectrum and is known to contain C, H, and O. Suggest several possible molecular formulas for nootkatone.

PROBLEM .....

12.3 By knowing the natural abundances of minor isotopes, it's possible to calculate the relative heights of  $M^{+}$  and  $(M + 1)^+$  peaks. If  $^{13}\text{C}$  has a natural abundance of 1.11%, what relative heights would you expect for  $M^{+}$  and  $(M + 1)^+$  peaks in the mass spectrum of benzene,  $\text{C}_6\text{H}_6$ ?

PROBLEM .....

12.4 The "nitrogen rule" of mass spectroscopy says that a compound containing an odd number of nitrogens has an odd-numbered molecular ion. Conversely, a compound containing an even number of nitrogens has an even-numbered  $M^{+}$  peak. Why should this be so?

PROBLEM .....

12.5 In light of the nitrogen rule mentioned in Problem 12.4, what is the molecular formula of pyridine,  $M^{+} = 79$ ?

## 12.7 Interpreting Mass Spectral Fragmentation Patterns

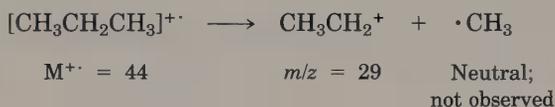
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Mass spectroscopy would be useful even if molecular weight and formula were the only information that could be obtained. In fact, though, we can get much more. The mass spectrum of a compound serves as a kind of

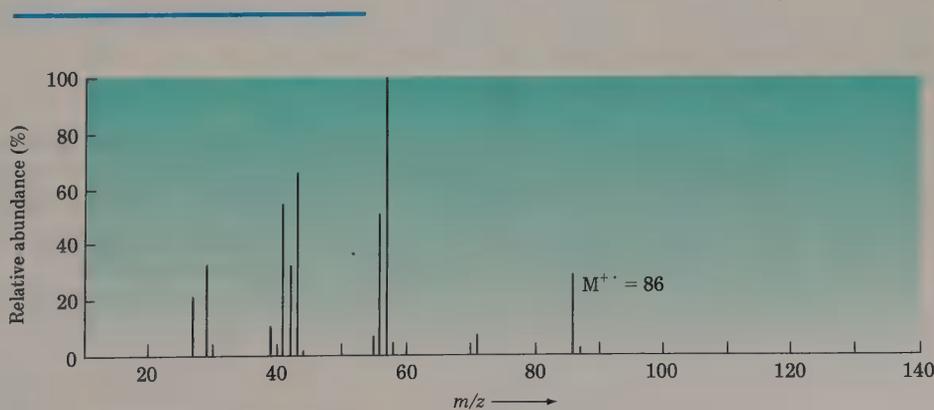
“molecular fingerprint.” Each organic molecule fragments in a unique way depending on its structure, and the chance that two compounds will have identical mass spectra is small. Thus, it’s sometimes possible to identify an unknown by computer-based matching of its mass spectrum to one of the more than 130,000 mass spectra recorded in a reference library established by the U.S. National Institutes of Health.

It’s also possible to derive structural information about a molecule by looking at its fragmentation pattern. Fragmentation occurs when the high-energy cation radical flies apart by spontaneous cleavage of a chemical bond. One of the two fragments retains the positive charge (a carbocation), and the other fragment is a neutral radical.

Not surprisingly, the positive charge usually remains with the fragment that is better able to stabilize it. In other words, the more stable carbocation is usually formed during fragmentation. For example, propane tends to fragment in such a way that the positive charge remains with the ethyl group rather than with the methyl, because the ethyl carbocation is more stable than the methyl carbocation (Section 6.11). Propane therefore has a base peak at  $m/z = 29$  and a barely detectable peak at  $m/z = 15$  (Figure 12.7).



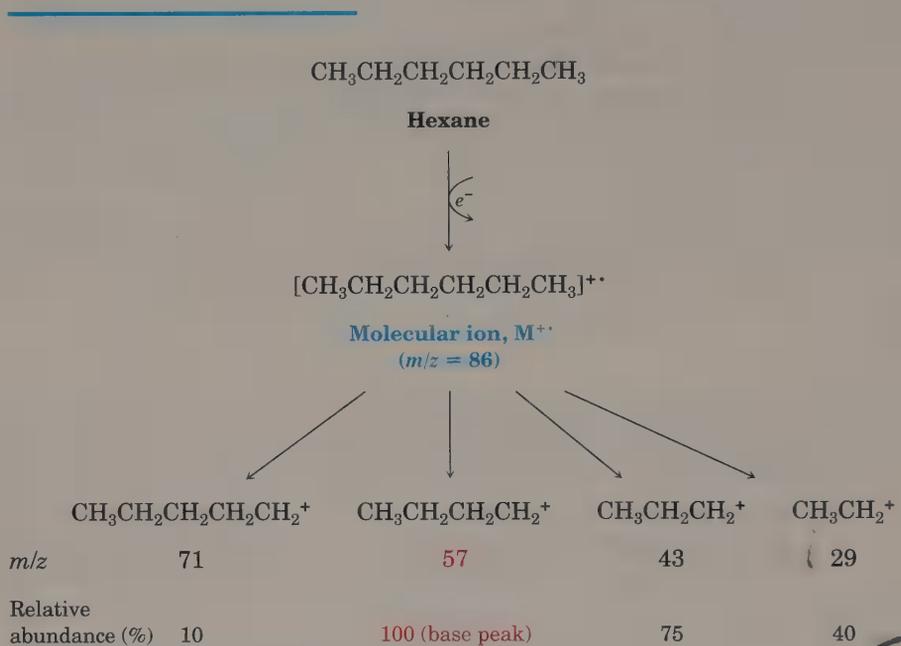
Since mass-spectral fragmentation patterns are usually complex, it’s often difficult to assign definite structures to fragment ions. Most hydrocarbons fragment in many ways, as the mass spectrum of hexane, a typical alkane, demonstrates (Figure 12.9). The hexane spectrum shows a moderately abundant molecular ion at  $m/z = 86$  and fragment ions at  $m/z = 71$ , 57, 43, and 29. Since all of the carbon–carbon bonds of hexane are



**Figure 12.9** Mass spectrum of hexane (C<sub>6</sub>H<sub>14</sub>; mol wt = 86). The base peak is at  $m/z = 57$ .

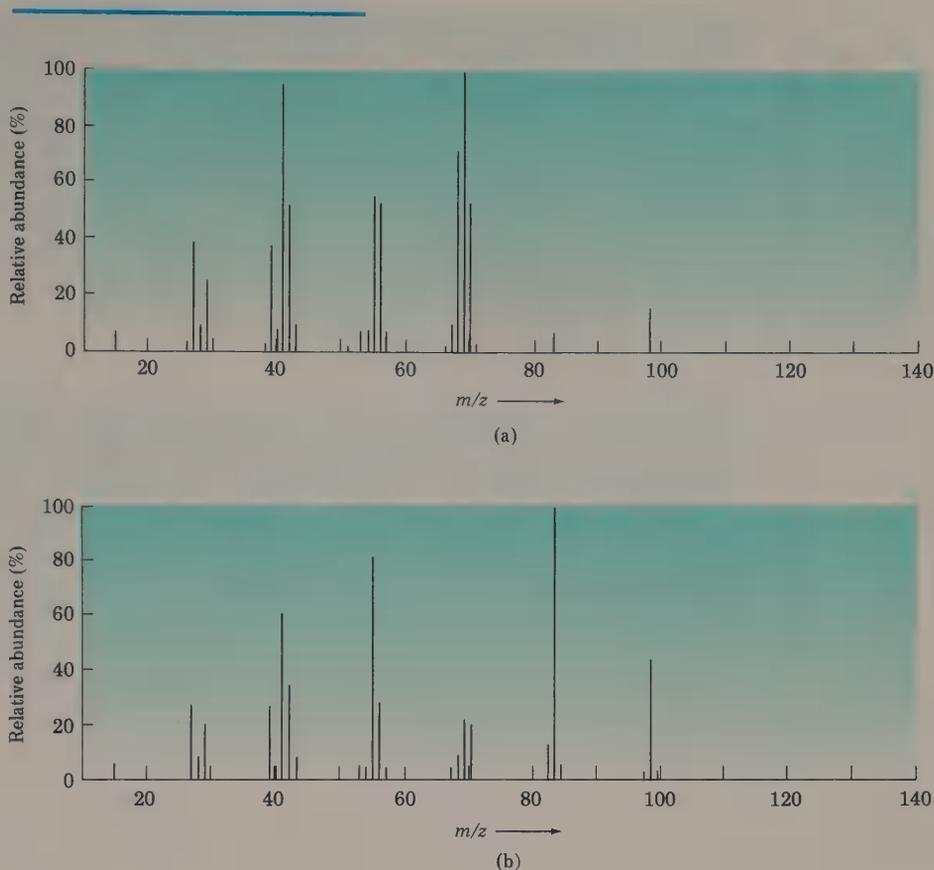
electronically similar, all of them break to a similar extent, giving rise to the observed ions.

Figure 12.10 shows how the hexane fragments might arise. The loss of a methyl radical from the hexane cation radical ( $M^{+\cdot} = 86$ ) gives rise to a fragment of mass 71; the loss of an ethyl radical accounts for a fragment of mass 57; the loss of a propyl radical accounts for a fragment of mass 43; and the loss of a butyl radical accounts for a fragment of mass 29. With skill and practice, chemists can learn to analyze the fragmentation patterns of unknown compounds and work backward to a structure that's compatible with the available data.



**Figure 12.10** Fragmentation of hexane in a mass spectrometer.

An example of how information from fragmentation patterns can be used to solve structural problems is given in Figure 12.11. Assume that we have two unlabeled bottles, A and B. One contains methylcyclohexane, the other contains ethylcyclopentane, and we need to distinguish between them. The mass spectra of both samples show molecular ions at  $M^{+\cdot} = 98$ , corresponding to  $\text{C}_7\text{H}_{14}$ , but the two spectra differ considerably in their fragmentation patterns. Sample B has its base peak at  $m/z = 83$ , corresponding to the loss of a  $\cdot\text{CH}_3$  group (15 mass units) from the molecular ion, but sample A has only a small peak at  $m/z = 83$ . Conversely, A has its base peak at  $m/z = 69$ , corresponding to the loss of a  $\cdot\text{CH}_2\text{CH}_3$  group (29 mass units), but B has a rather small peak at  $m/z = 69$ . We can therefore be reasonably certain that B is methylcyclohexane and A is ethylcyclopentane.



**Figure 12.11** Mass spectra of unknown samples A and B.

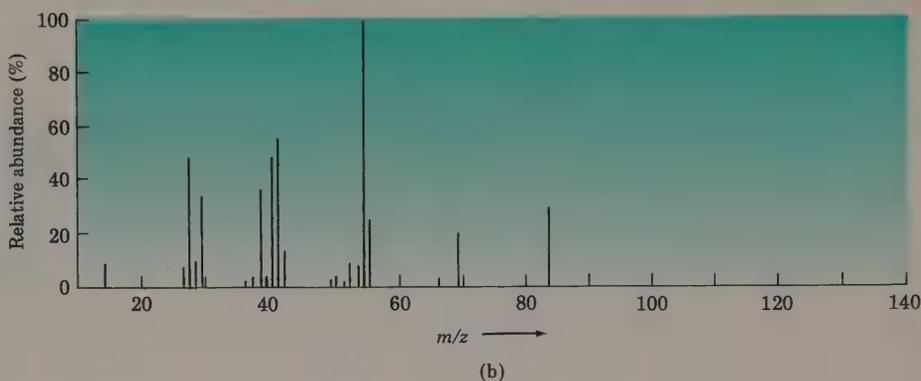
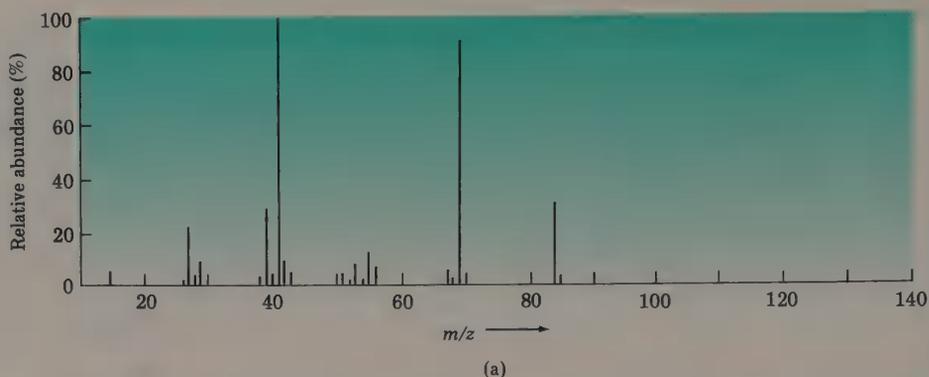
This example is a simple one, but the principles used are broadly applicable for organic structure determination by mass spectroscopy. As we'll see in later chapters, specific functional-group families such as alcohols, ketones, aldehydes, and amines often show specific kinds of mass spectral fragmentations that can be interpreted to provide structural information.

**PROBLEM** .....

- 12.6** The mass spectrum of 2,2-dimethylpropane (Figure 12.8) shows a base peak at  $m/z = 57$ . What molecular formula does this correspond to? Can you suggest a structure for the  $m/z = 57$  fragment ion?

**PROBLEM** .....

- 12.7** Two mass spectra are shown on the following page. One spectrum corresponds to 2-methyl-2-pentene; the other, to 2-hexene. Which do you think is which? Explain your choices.

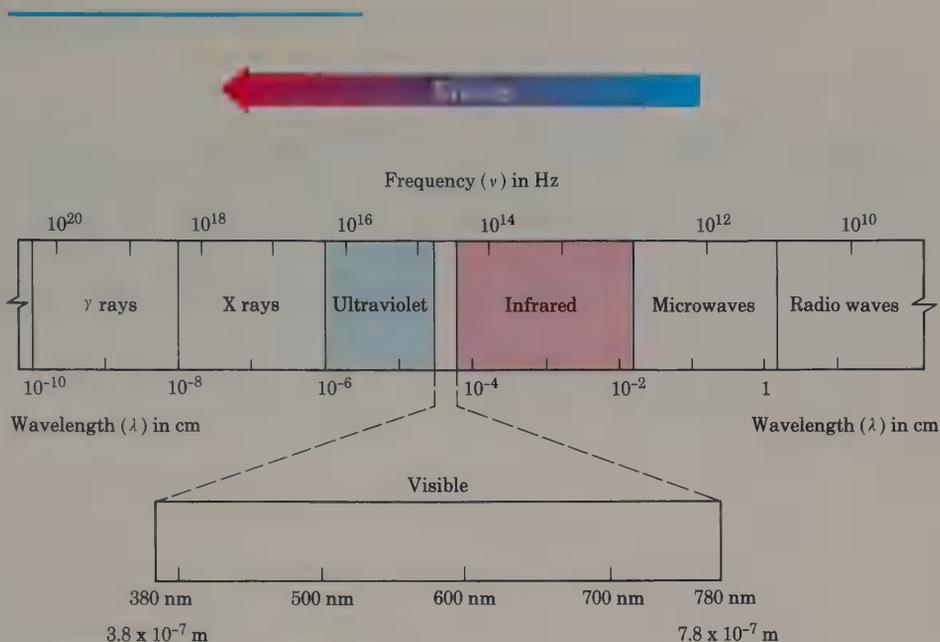


## 12.8 Spectroscopy and the Electromagnetic Spectrum

Infrared, ultraviolet, and nuclear magnetic resonance spectroscopies differ from mass spectroscopy in that they involve the interaction of molecules with electromagnetic energy rather than with a high-energy electron beam. Before beginning a study of these techniques, we need to look into the nature of radiant energy and the electromagnetic spectrum.

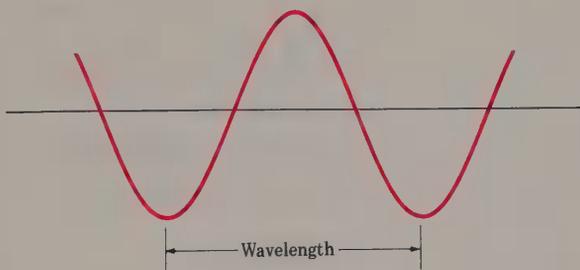
Visible light, X rays, microwaves, radio waves, and so forth are all different kinds of **electromagnetic radiation**. Collectively, they make up the **electromagnetic spectrum**, shown in Figure 12.12. As indicated, the electromagnetic spectrum is arbitrarily divided into various regions, with the familiar visible region accounting for only a small portion of the overall spectrum (from  $3.8 \times 10^{-5}$  cm to  $7.8 \times 10^{-5}$  cm in wavelength). The visible region is flanked by the infrared and ultraviolet regions.

Electromagnetic radiation can be thought of as having dual behavior. In some respects it has the properties of a particle (called a *photon*), yet in other respects it behaves as a wave traveling at the speed of light. Electromagnetic waves can be described by their wavelength ( $\lambda$ ; Greek lambda) and frequency ( $\nu$ ; Greek nu). The **wavelength** is simply the length of one complete wave cycle from trough to trough; the **frequency** is the number



**Figure 12.12** The electromagnetic spectrum.

of wave cycles that travel past a fixed point in a certain unit of time (usually given in cycles per second, or **hertz, Hz**). Wavelength and frequency are inversely related by the equation  $\lambda = c/\nu$ , where  $\lambda$  is measured in centimeters,  $c$  is the speed of light ( $3 \times 10^{10}$  cm/sec), and  $\nu$  is measured in hertz.



Electromagnetic energy is transmitted only in discrete energy bundles, called **quanta**. The amount of energy corresponding to 1 quantum of energy (or 1 photon) of a given frequency is expressed by the equation:

$$\epsilon = h\nu = hc/\lambda$$

where  $\epsilon$  = The energy of 1 photon (1 quantum)

$h$  = Planck's constant ( $6.62 \times 10^{-34}$  joule sec =  $1.58 \times 10^{-34}$  cal sec)

$\nu$  = Frequency in hertz

$\lambda$  = Wavelength in centimeters

$c$  = Speed of light ( $3 \times 10^{10}$  cm/sec)

This equation says that the amount of energy in a photon varies *directly* with its frequency, but *inversely* with its wavelength. High frequencies and short wavelengths correspond to high-energy radiation such as gamma rays; low frequencies and long wavelengths correspond to low-energy radiation such as radio waves. If we multiply  $\varepsilon$  by Avogadro's number,  $N$ , and convert to kilocalories per mole, we arrive at the same equation expressed in units familiar to organic chemists:

$$E = \frac{Nhc}{\lambda} = \frac{2.86 \times 10^{-3} \text{ kcal/mol}}{\lambda \text{ (cm)}}$$

where  $E$  represents the energy of Avogadro's number (a "mole") of photons of wavelength  $\lambda$

When an organic compound is exposed to electromagnetic radiation, energy of certain wavelengths is absorbed, but energy of other wavelengths passes through. Whether the light energy is absorbed or not absorbed depends both on the structure of the sample compound and on the wavelength of the radiation. If we irradiate the sample with energy of many different wavelengths and determine which wavelengths are absorbed and which pass through, we can determine the **absorption spectrum** of the compound.

Absorption spectra are usually displayed on graphs that plot wavelength versus amount of radiation transmitted. For example, the absorption spectrum of ethyl alcohol using infrared radiation is shown in Figure 12.13. The horizontal axis records the wavelength in micrometers ( $\mu\text{m}$ ), and the vertical axis records the intensity of the various energy absorptions in % transmittance. The baseline corresponding to 0% absorption (or 100% transmittance) runs along the top of the chart, and a downward spike means that energy absorption has occurred at that wavelength.

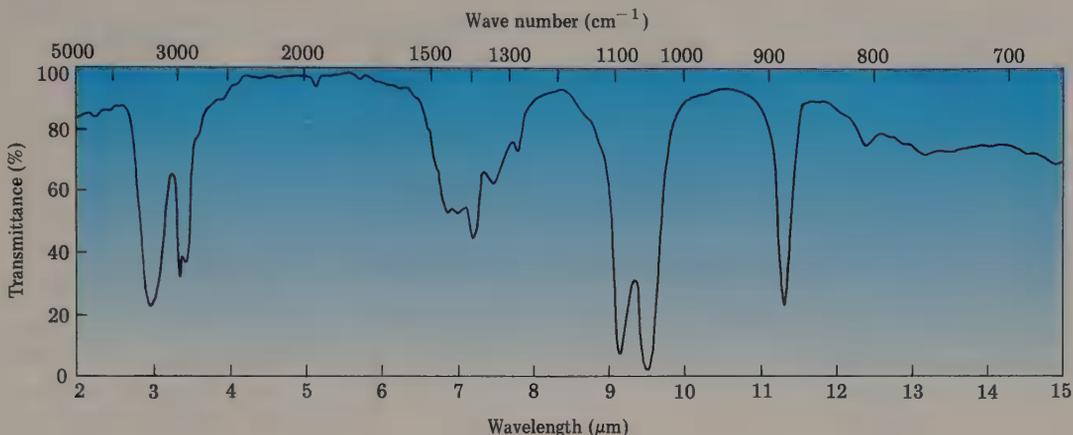
Since a molecule gains energy when it absorbs radiation, the energy must be distributed over the molecule in some way. For example, energy absorption might result in increased molecular motions, causing bonds to stretch, bend, or rotate. Alternatively, energy absorption might cause electrons to be excited from a low-energy orbital to a higher one. Different radiation frequencies affect molecules in different ways, but each can provide structural information if we learn to interpret the results properly.

There are numerous kinds of spectroscopy, depending on which region of the electromagnetic spectrum is used. We'll look closely at two, infrared spectroscopy and nuclear magnetic resonance spectroscopy, and have a brief introduction to a third, ultraviolet spectroscopy. Let's begin by seeing what happens when an organic sample absorbs infrared energy.

#### PRACTICE PROBLEM.....

Which is higher in energy, FM radio waves with a frequency of  $1.015 \times 10^8$  Hz (101.5 MHz) or visible light with a frequency of  $5 \times 10^{14}$  Hz?

**Solution** The equation  $\varepsilon = h\nu$  says that energy increases as frequency increases. Thus, visible light is higher in energy than radio waves.



**Figure 12.13** The infrared spectrum of ethyl alcohol. A transmittance of 100% means that all the energy is passing through the sample, whereas a lower transmittance means that some energy is being absorbed. Thus, each downward spike corresponds to an energy absorption.

**PROBLEM** .....

- 12.8** How does the energy of infrared radiation with  $\lambda = 10^{-4}$  cm compare with that of an X ray with  $\lambda = 3 \times 10^{-7}$  cm?

**PROBLEM** .....

- 12.9** Which is higher in energy, radiation with  $\nu = 4 \times 10^9$  Hz or radiation with  $\lambda = 9 \times 10^{-4}$  cm?

**PROBLEM** .....

- 12.10** It's useful to develop an intuitive feeling for the amounts of energy that correspond to different parts of the electromagnetic spectrum. Using the relationships

$$E = \frac{2.86 \times 10^{-3} \text{ kcal/mol}}{\lambda \text{ (cm)}} \quad \text{and} \quad \nu = \frac{c}{\lambda}$$

calculate the energies of the following:

- A gamma ray with  $\lambda = 5 \times 10^{-9}$  cm
- An X ray with  $\lambda = 3 \times 10^{-7}$  cm
- Ultraviolet light with  $\nu = 6 \times 10^{15}$  Hz
- Visible light with  $\nu = 7 \times 10^{14}$  Hz
- Infrared radiation with  $\lambda = 2 \times 10^{-3}$  cm

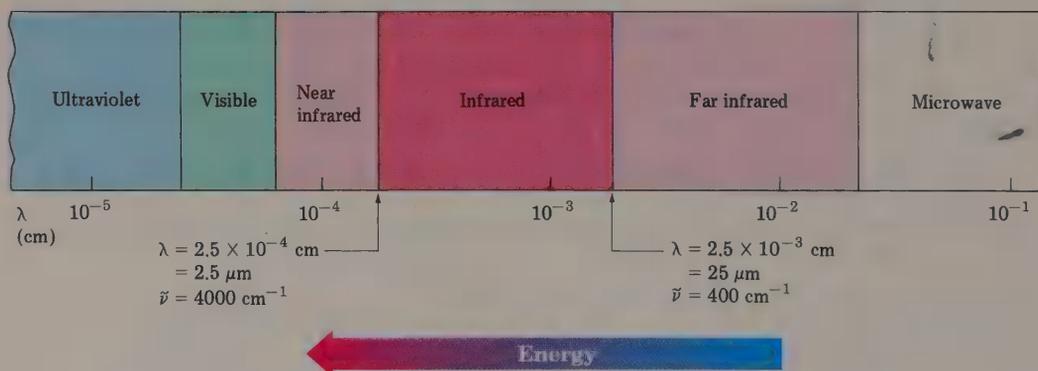
- (f) Microwave radiation with  $\nu = 10^{11}$  Hz  
 (g) Radio waves from the fat one, station KFAT in Gilroy, California, 102.5 MHz FM.

## 12.9 Infrared Spectroscopy of Organic Molecules

The infrared (IR) region of the electromagnetic spectrum covers the range from just above the visible ( $7.8 \times 10^{-5}$  cm) to approximately  $10^{-2}$  cm, but only the midportion from  $2.5 \times 10^{-3}$  cm to  $2.5 \times 10^{-4}$  cm interests organic chemists (Figure 12.14). Specific wavelengths within the IR region are usually given in micrometers ( $1 \mu\text{m} = 10^{-4}$  cm), and frequencies are expressed in **wave numbers** ( $\tilde{\nu}$ ) rather than in hertz. The wave number, expressed in units of  $\text{cm}^{-1}$ , is simply the reciprocal of the wavelength in centimeters.

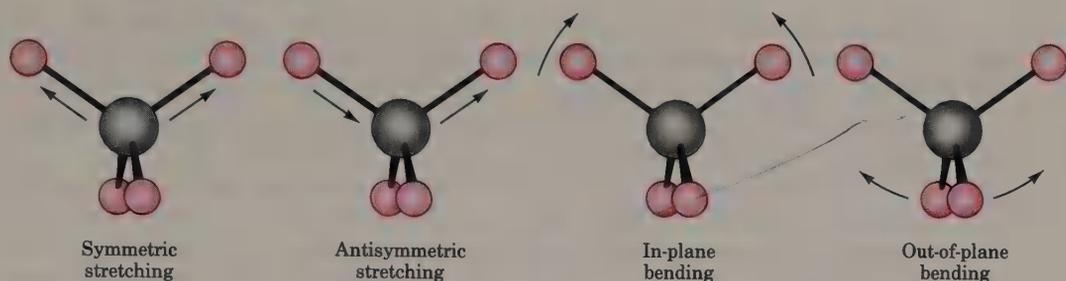
$$\text{Wave number } (\tilde{\nu}) = 1/\lambda \text{ (cm)}$$

Using the equation  $E = (2.86 \times 10^{-3} \text{ kcal/mol})/\lambda$ , we can calculate that the energy levels of infrared radiation range from 1.13 to 11.3 kcal/mol (4.73–47.3 kJ/mol).



**Figure 12.14** The infrared region of the electromagnetic spectrum.

Why does an organic molecule absorb some wavelengths of infrared radiation but not others? All molecules have a certain amount of energy distributed throughout their structure causing bonds to stretch and bend, atoms to wag and rock, and other molecular vibrations to occur. Some of the kinds of allowed vibrations are shown:



The amount of energy a molecule contains is not continuously variable but is *quantized*. That is, a molecule can stretch, bend, or vibrate only at specific frequencies corresponding to specific energy levels. Take bond lengths for example. Although we usually cite bond lengths as if they were fixed, the numbers we give are actually averages. In reality, bonds are constantly stretching and bending, lengthening and contracting. Thus, a typical C–H bond with an average bond length of 1.10 Å is actually vibrating back and forth at a specific frequency, alternately stretching and compressing as if there were a spring connecting the two atoms. When the molecule is irradiated with electromagnetic radiation, *the vibrating bond will absorb energy if the frequencies of the light and the vibration are the same*.

When a molecule absorbs infrared radiation, the molecular vibration with a frequency matching that of the light increases in amplitude. In other words, the “spring” connecting the two atoms stretches and compresses a bit further. Since each light frequency absorbed by the molecule corresponds to a specific molecular motion, we can see what kinds of molecular motions a sample has by measuring its infrared spectrum. By then working backward and interpreting the infrared spectrum, we can find out what kinds of bonds (functional groups) are present in the molecule.

**IR spectrum** → **What molecular motions?** →

**What functional groups?**

**PROBLEM** .....

- 12.11** Some chemists prefer to express IR absorptions in micrometers, whereas others prefer wave numbers. To converse with both groups, it's useful to be able to interconvert the two systems of measurement rapidly. Do the following conversions:
- (a)  $3.1 \mu\text{m}$  to  $\text{cm}^{-1}$       (b)  $5.85 \mu\text{m}$  to  $\text{cm}^{-1}$       (c)  $2250 \text{cm}^{-1}$  to  $\mu\text{m}$   
 (d)  $970 \text{cm}^{-1}$  to  $\mu\text{m}$
- .....

## 12.10 Interpreting Infrared Spectra

The full interpretation of an IR spectrum is difficult because most organic molecules are so large that there are dozens or hundreds of different possible bond stretching and bending motions. Thus, an IR spectrum contains dozens or hundreds of absorptions. In one sense this complexity is valuable, since an infrared spectrum serves as a unique fingerprint of a specific compound.

In fact, the complex region of the infrared spectrum below  $1500\text{ cm}^{-1}$  is called the **fingerpr**int** region**; if two samples have identical infrared spectra, the compounds are almost certainly identical. For structural purposes, however, the multitude of absorptions present in an infrared spectrum makes full interpretation difficult.

Fortunately, we don't need to interpret a spectrum fully to get useful structural information. *Most functional groups have characteristic infrared absorptions that don't change much from one compound to another.* The C=O absorption of ketones is almost always in the range  $1690\text{--}1750\text{ cm}^{-1}$ ; the O-H absorption of alcohols is almost always in the range  $3200\text{--}3600\text{ cm}^{-1}$ ; and the C=C absorption of alkenes is almost always in the range  $1640\text{--}1680\text{ cm}^{-1}$ . By learning to recognize where characteristic functional-group absorptions occur, it's possible to obtain valuable structural information from IR spectra. Table 12.1 lists the characteristic IR absorptions of some common functional groups.

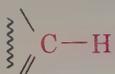
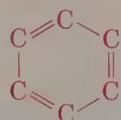
Look at the IR spectra of hexane, 1-hexene, and 1-hexyne in Figure 12.15 (page 432) to see an example of how infrared spectroscopy can be used. Although all three IR spectra contain many peaks, there are characteristic absorptions of the C=C and C≡C functional groups that allow the three compounds to be distinguished readily. Thus, 1-hexene shows a characteristic carbon-carbon double-bond peak at  $1660\text{ cm}^{-1}$  and a vinylic C-H bond peak at  $3100\text{ cm}^{-1}$ , whereas 1-hexyne has a carbon-carbon triple-bond absorption at  $2100\text{ cm}^{-1}$  and a terminal alkyne ≡C-H bond absorption at  $3300\text{ cm}^{-1}$ .

It helps in remembering the position of specific IR absorptions to divide the infrared region from  $4000$  to  $200\text{ cm}^{-1}$  into four parts, as shown in Figure 12.16 (page 433):

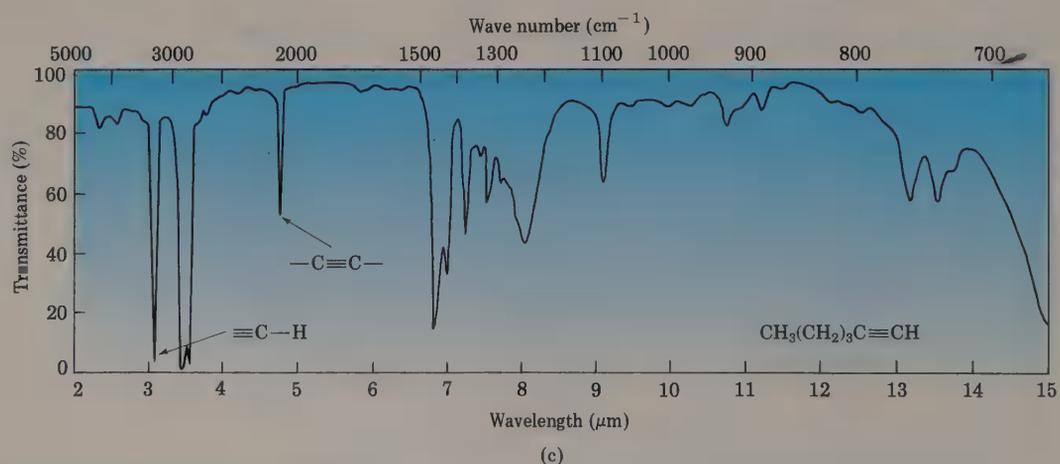
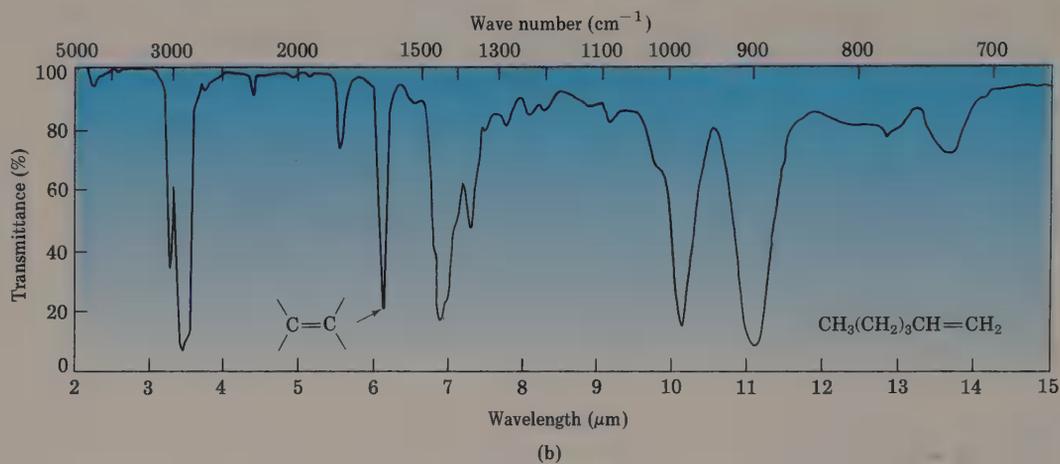
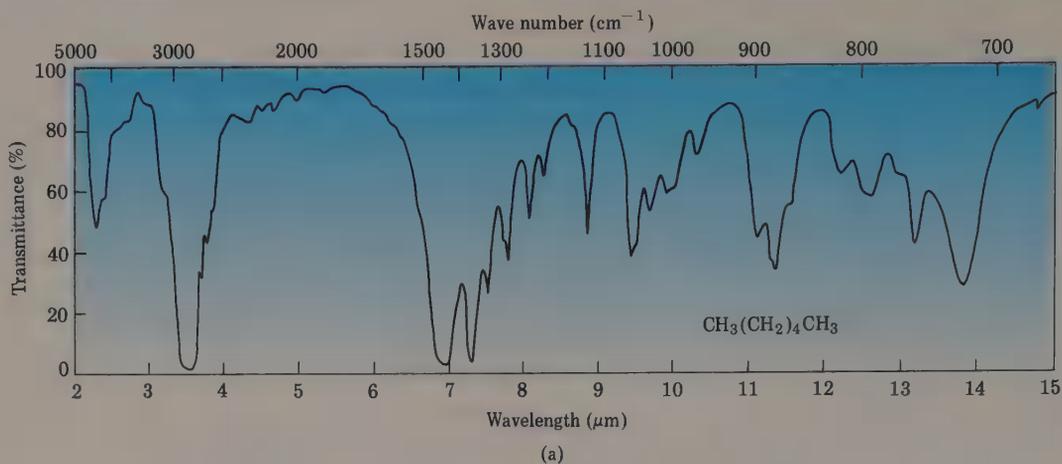
1. The region from  $4000$  to  $2500\text{ cm}^{-1}$  corresponds to absorptions caused by N-H, C-H, and O-H single-bond stretching motions. The N-H and O-H bonds absorb in the  $3300\text{--}3600\text{ cm}^{-1}$  range, whereas C-H bond stretching occurs near  $3000\text{ cm}^{-1}$ .
2. The region from  $2500$  to  $2000\text{ cm}^{-1}$  is where triple-bond stretching occurs. Both nitriles (RC≡N) and alkynes show peaks here.
3. The region from  $2000$  to  $1500\text{ cm}^{-1}$  is where double bonds of all kinds (C=O, C=N, and C=C) absorb. Carbonyl groups generally absorb in the range from  $1670$  to  $1780\text{ cm}^{-1}$ , and alkene stretching normally occurs in a narrow range from  $1640$  to  $1680\text{ cm}^{-1}$ .
4. The region below  $1500\text{ cm}^{-1}$  is the fingerprint portion of the IR range. A large number of absorptions due to a variety of C-C, C-O, C-N, and C-X single-bond vibrations occur here.

Why do different functional groups absorb where they do? The best analogy is that of two weights (atoms) connected by a spring (a bond). Short, strong bonds vibrate at a higher frequency and with lower wavelength than long, weak bonds, just as a short, strong spring vibrates faster than a long, weak spring. Thus, triple bonds absorb at a higher frequency than double bonds, which in turn absorb higher than single bonds. In addition, springs connecting small weights vibrate faster than springs connecting large weights. Thus, C-H, O-H, and N-H bonds vibrate faster than bonds between heavier C, O, and N atoms.

TABLE 12.1 Characteristic Infrared Absorptions of Some Functional Groups

Functional group class	Band position ( $\text{cm}^{-1}$ )	Intensity of absorption
Alkanes, alkyl groups C—H	2850–2960	Medium to strong
Alkenes =C—H C=C	3020–3100 1650–1670	Medium Medium
Alkynes $\equiv\text{C—H}$ —C $\equiv\text{C—}$	3300 2100–2260	Strong Medium
Alkyl halides C—Cl C—Br C—I	600–800 500–600 500	Strong Strong Strong
Alcohols O—H C—O	3400–3640 1050–1150	Strong, broad Strong
Aromatics 	3030	Medium
	1600, 1500	Strong
Amines N—H C—N	3310–3500 1030, 1230	Medium Medium
Carbonyl compounds <sup>a</sup> C=O	1670–1780	Strong
Carboxylic acids O—H	2500–3100	Strong, very broad
Nitriles C $\equiv\text{N}$	2210–2260	Medium
Nitro compounds NO <sub>2</sub>	1540	Strong

<sup>a</sup>Acids, esters, aldehydes, and ketones.



**Figure 12.15** Infrared spectra of (a) hexane, (b) 1-hexene, and (c) 1-hexyne. Spectra like these are easily obtained on small samples in a few minutes using commercially available instruments.

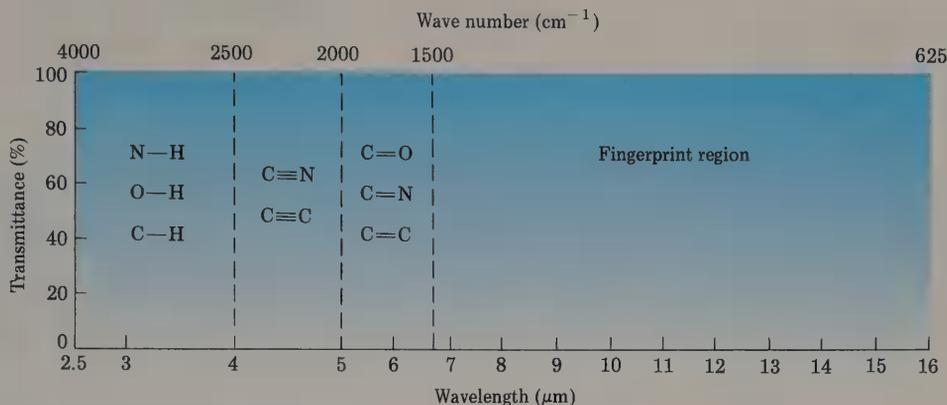


Figure 12.16 Regions in the infrared spectrum.

PROBLEM .....

12.12 Refer to Table 12.1, and make educated guesses as to what functional groups these molecules might contain:

- A compound with a strong absorption at  $1710\text{ cm}^{-1}$
- A compound with a strong absorption at  $1540\text{ cm}^{-1}$
- A compound with strong absorptions at  $1720\text{ cm}^{-1}$  and at  $2500\text{--}3100\text{ cm}^{-1}$

PROBLEM .....

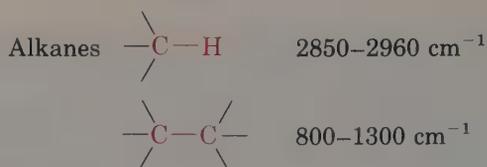
12.13 How might you use IR spectroscopy to help distinguish between these pairs of isomers?

- $\text{CH}_3\text{CH}_2\text{OH}$  and  $\text{CH}_3\text{OCH}_3$
- Cyclohexane and 1-hexene
- $\text{CH}_3\text{CH}_2\text{COOH}$  and  $\text{HOCH}_2\text{CH}_2\text{CHO}$

## 12.11 Infrared Spectra of Hydrocarbons

### ALKANES

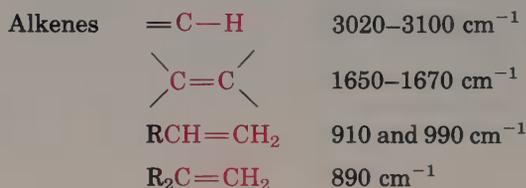
The infrared spectrum of an alkane is rather uninformative, since no functional groups are present and all absorptions are due to C–H and C–C bond stretching and bending. Alkane C–H bonds always show a strong absorption from  $2850$  to  $2960\text{ cm}^{-1}$ , and saturated C–C bonds show a number of absorptions in the  $800\text{--}1300\text{ cm}^{-1}$  range. Since most organic compounds contain saturated alkane-like portions, most organic compounds have these characteristic IR peaks. These C–H and C–C peaks are clearly visible in the three spectra shown in Figure 12.15.



## ALKENES

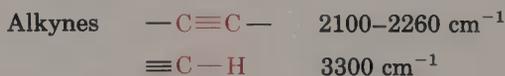
Alkenes show several characteristic stretching absorptions that can be used for structural identification. For example, vinylic  $\text{=C-H}$  bonds absorb from 3020 to 3100  $\text{cm}^{-1}$  and alkene  $\text{C=C}$  bonds usually show an absorption near 1650  $\text{cm}^{-1}$ , although in some cases this can be rather weak and difficult to see clearly. Both  $\text{=C-H}$  and  $\text{C=C}$  absorptions are diagnostic for alkenes, as the 1-hexene spectrum in Figure 12.15 shows.

Mono- and disubstituted alkenes have  $\text{=C-H}$  bonds with specific out-of-plane bending absorptions in the 700–1000  $\text{cm}^{-1}$  range, thereby allowing the substitution pattern on a double bond to be determined. Monosubstituted alkenes such as 1-hexene show strong characteristic peaks at 910 and 990  $\text{cm}^{-1}$ , and 2,2-disubstituted alkenes ( $\text{R}_2\text{C=CH}_2$ ) have an intense band at 890  $\text{cm}^{-1}$ .



## ALKYNES

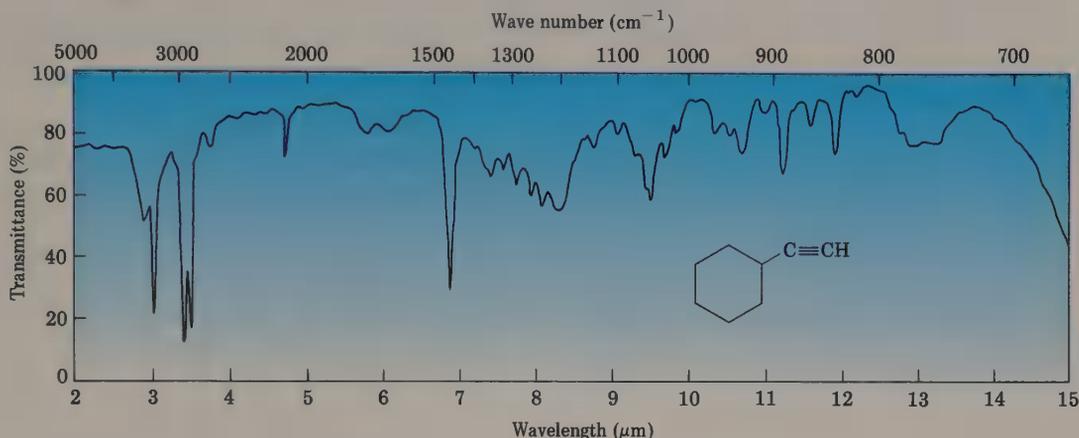
Alkynes show a  $\text{C}\equiv\text{C}$  stretching absorption at 2100–2260  $\text{cm}^{-1}$ , a band that is much more intense for terminal alkynes than for internal alkynes. In fact, symmetrically substituted triple bonds like that in 3-hexyne show no absorption at all, for reasons we won't go into. Terminal alkynes such as 1-hexyne also have a characteristic  $\equiv\text{C-H}$  stretch at 3300  $\text{cm}^{-1}$ . This band is diagnostic for terminal alkynes, since it is fairly intense and quite sharp.



One other point about infrared spectroscopy: It's also possible to derive much structural information from an IR spectrum by noticing which absorptions are *not* present. If the spectrum of an unknown has no absorptions at 3300 and 2150  $\text{cm}^{-1}$ , the unknown isn't a terminal alkyne; if the spectrum has no absorption near 3400  $\text{cm}^{-1}$ , the compound isn't an alcohol; and so on.

## PROBLEM.....

12.14 Shown here is the infrared spectrum of ethynylcyclohexane. What absorption bands can you identify?

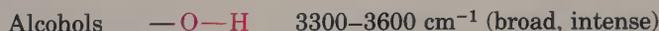


## 12.12 Infrared Spectra of Other Functional Groups

As each functional group is discussed in future chapters, the spectroscopic behavior of that group will be described. For the present, though, it's worthwhile to point out some distinguishing features of the more important functional groups.

### ALCOHOLS

The O-H functional group of alcohols is easy to spot in the IR. Alcohols have a characteristic band in the range  $3300\text{--}3600\text{ cm}^{-1}$  that is usually fairly broad and intense. If present, it's hard to miss this band or to confuse it with anything else.



### AMINES

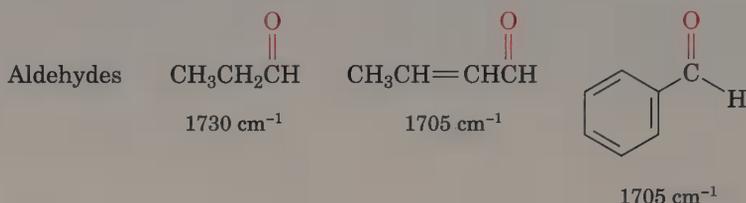
The N-H functional group of amines is also easy to spot in the IR, with a characteristic absorption in the  $3300\text{--}3500\text{ cm}^{-1}$  range. Although alcohols also absorb in this range, an amine absorption is much sharper and less intense than a hydroxyl band.



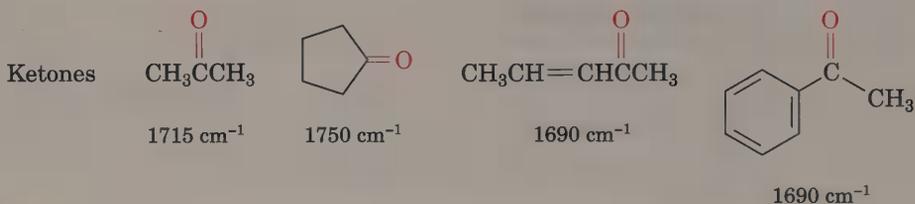
## CARBONYL COMPOUNDS

Carbonyl functional groups are the easiest to identify of all IR absorptions, with a sharp, intense peak in the range  $1670\text{--}1780\text{ cm}^{-1}$ . Most important, the exact position of absorption within the range can often be used to identify the exact kind of carbonyl functional group—aldehyde, ketone, ester, and so forth.

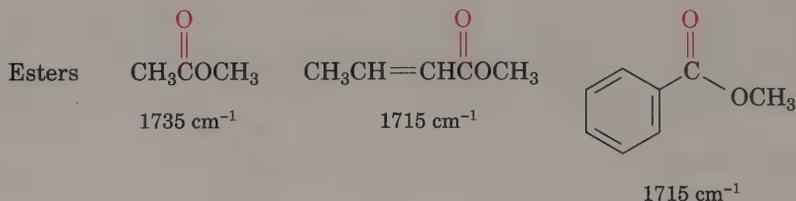
**Aldehydes** Saturated aldehydes absorb at  $1730\text{ cm}^{-1}$ , whereas aldehydes next to either a double bond or an aromatic ring absorb at  $1705\text{ cm}^{-1}$ .



**Ketones** Open-chain ketones and six-membered-ring cyclic ketones absorb at  $1715\text{ cm}^{-1}$ ; five-membered-ring ketones absorb at  $1750\text{ cm}^{-1}$ ; and ketones next to a double bond or an aromatic ring absorb at  $1690\text{ cm}^{-1}$ .

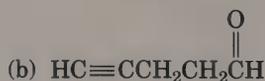
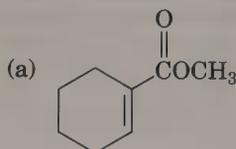


**Esters** Saturated esters absorb at  $1735\text{ cm}^{-1}$ , and esters next to either an aromatic ring or a double bond absorb at  $1715\text{ cm}^{-1}$ .



## PROBLEM .....

12.15 Make guesses about where the following compounds might have characteristic IR absorptions:



.....

## 12.13 Summary and Key Words

Product purification and structure elucidation are two of the most difficult tasks facing laboratory chemists. If possible, purification is done by a direct method such as **crystallization** or **distillation**. Often, however, some form of **chromatography** must be used. The impure sample is applied to one end of a column containing an inert support material (the **stationary phase**), and the sample is carried along by a **mobile phase**, either gas or liquid. Separation of mixtures is effected because different compounds adsorb to the stationary phase in differing degrees and therefore pass along the column at different rates.

Once a reaction product has been purified, its structure must be determined using various spectroscopic methods such as mass spectroscopy and infrared spectroscopy. **Mass spectroscopy** gives information about the molecular weight and formula of unknown samples; **infrared spectroscopy** gives information about the functional groups present.

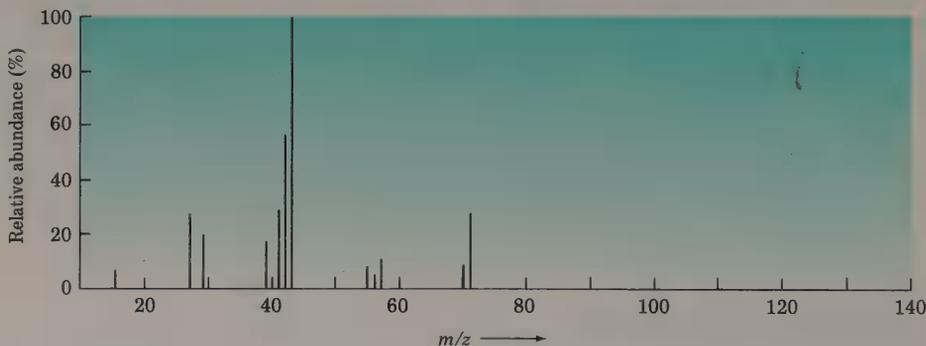
In mass spectroscopy, molecules are first ionized by collision with a high-energy electron beam. The ions then fragment into smaller pieces, which are magnetically sorted according to their mass-to-charge ratio ( $m/z$ ). The ionized sample molecule is called the **molecular ion**,  $M^{+}$ , and measurement of its mass gives the molecular weight of the sample. Structural clues about unknown samples can be obtained by interpreting the **fragmentation pattern** of the molecular ion. Mass spectral fragmentations are usually complex, however, and interpretation is often difficult.

Infrared spectroscopy involves the interaction of a molecule with **electromagnetic radiation**. When an organic molecule is irradiated with infrared light, certain frequencies of light are absorbed by the molecule and others are not absorbed. The frequencies absorbed correspond to the amounts of energy needed to increase the amplitude of specific molecular vibrations such as bond stretchings and bendings. Every functional group has a characteristic set of infrared absorptions. For example, the terminal alkyne  $\equiv C-H$  bond absorbs infrared radiation of  $3300\text{ cm}^{-1}$  frequency, and the alkene  $C=C$  bond absorbs in the range  $1650\text{--}1670\text{ cm}^{-1}$ . By observing which frequencies of infrared radiation are absorbed by a molecule and which are not, it's possible to determine what functional groups a molecule contains.

### ADDITIONAL PROBLEMS

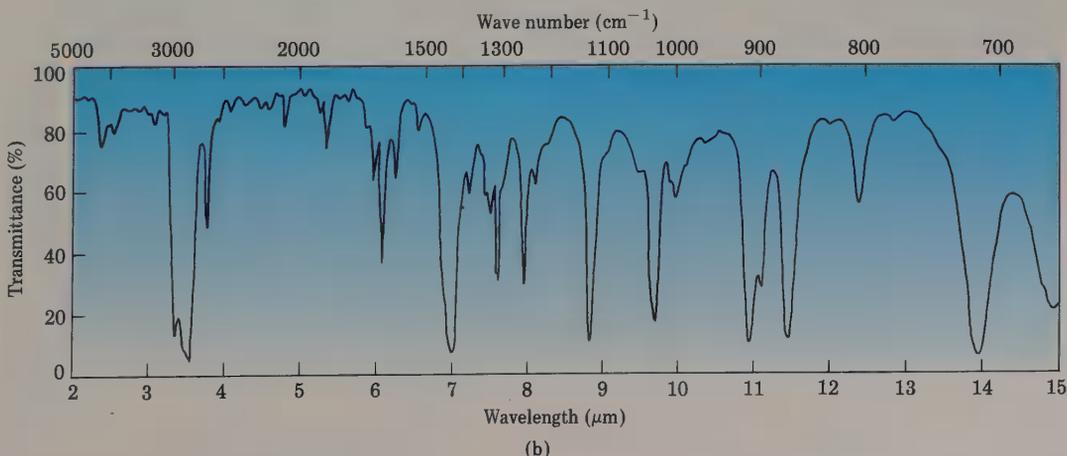
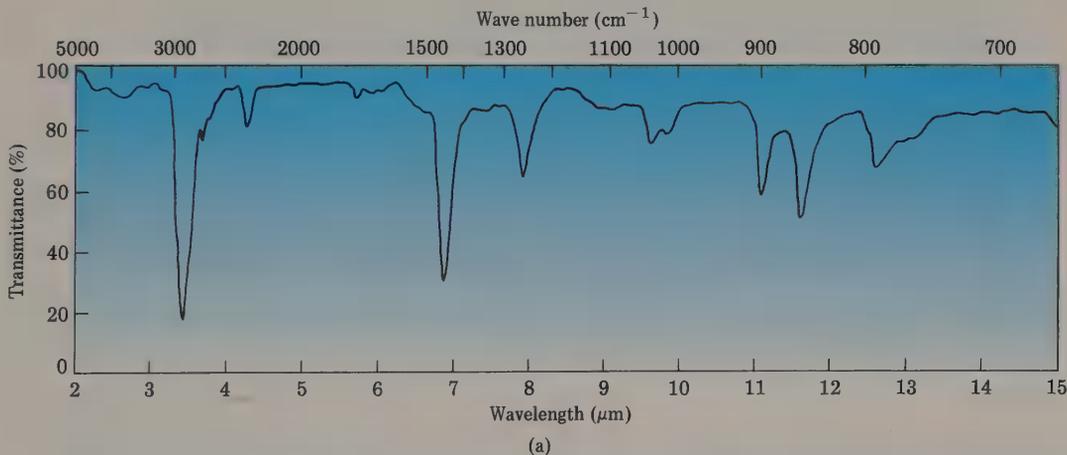
- 12.16 Write as many molecular formulas as you can for hydrocarbons that show the following molecular ions in their mass spectra:  
 (a)  $M^{+} = 64$       (b)  $M^{+} = 186$       (c)  $M^{+} = 158$       (d)  $M^{+} = 220$
- 12.17 In Section 6.2, we calculated the degree of unsaturation of molecules according to their molecular formulas. Write the molecular formulas of all hydrocarbons corresponding to the following molecular ions. How many degrees of unsaturation (double bonds and/or rings) are indicated by each formula?  
 (a)  $M^{+} = 86$       (b)  $M^{+} = 110$       (c)  $M^{+} = 146$       (d)  $M^{+} = 190$
- 12.18 Draw the structure of a molecule that is consistent with the mass spectral data in each example:  
 (a) A hydrocarbon with  $M^{+} = 132$       (b) A hydrocarbon with  $M^{+} = 166$   
 (c) A hydrocarbon with  $M^{+} = 84$

- 12.19 Write as many molecular formulas as you can for compounds that show the following molecular ions in their mass spectra. Assume that the elements C, H, N, and O might be present.  
 (a)  $M^{+\cdot} = 74$  (b)  $M^{+\cdot} = 131$
- 12.20 Camphor, a saturated monoketone from the Asian camphor tree, is used, among other things, as a moth repellent and as a constituent of embalming fluid. If camphor has  $M^{+\cdot} = 152$  in its mass spectrum, what is a reasonable molecular formula? How many rings does camphor have?
- 12.21 Nicotine is a diamino compound that can be isolated from dried tobacco leaves. Nicotine has two rings and shows  $M^{+\cdot} = 162$  in its mass spectrum. Propose a molecular formula for nicotine, and calculate the number of double bonds present. (There is no oxygen in nicotine.)
- 12.22 Halogenated compounds are particularly easy to identify by their mass spectra because both chlorine and bromine occur naturally as mixtures of two abundant isotopes. Chlorine occurs as  $^{35}\text{Cl}$  (75.5%) and as  $^{37}\text{Cl}$  (24.5%), and bromine occurs as  $^{79}\text{Br}$  (50.5%) and  $^{81}\text{Br}$  (49.5%). At what masses do the molecular ion(s) occur for these formulas? What are the relative percentages of each molecular ion?  
 (a) Bromomethane,  $\text{CH}_3\text{Br}$  (b) 1-Chlorohexane,  $\text{C}_6\text{H}_{13}\text{Cl}$
- 12.23 Molecular ions can be particularly complex for polyhalogenated compounds. Taking the natural abundance of Cl into account (Problem 12.22), calculate the masses of the molecular ions of the following formulas. What are the relative percentages of each ion?  
 (a) Chloroform,  $\text{CHCl}_3$  (b) Freon 12,  $\text{CF}_2\text{Cl}_2$  (fluorine occurs only as  $^{19}\text{F}$ )
- 12.24 2-Methylpentane ( $\text{C}_6\text{H}_{14}$ , mol wt 86) has the mass spectrum shown here. What peak represents  $M^{+\cdot}$ ? Which is the base peak? Propose structures for fragment ions of  $m/z = 71$ , 57, 43, and 29. Suggest a reason for the base peak's having the mass it does.

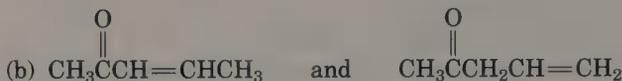
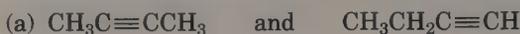


- 12.25 The combined gas chromatograph/mass spectrometer is a sophisticated instrument that uses a mass spectrometer to detect compounds as they elute from a gas chromatograph. This technique allows one to inject a mixture of compounds onto a gas chromatography column and to determine automatically the mass spectrum of each compound present in the mixture. Assume that you are in the laboratory carrying out the catalytic hydrogenation of cyclohexene to cyclohexane. How could you use a gas chromatograph/mass spectrometer to determine when the reaction was complete?
- 12.26 Convert the following infrared absorption values from micrometers to wave numbers:  
 (a) An alcohol,  $2.98 \mu\text{m}$  (b) An ester,  $5.81 \mu\text{m}$  (c) A nitrile,  $4.93 \mu\text{m}$

- 12.27** Convert the following infrared absorption values from wave numbers to micrometers:  
 (a) A cyclopentanone,  $1755\text{ cm}^{-1}$                       (b) An amine,  $3250\text{ cm}^{-1}$   
 (c) An aldehyde,  $1725\text{ cm}^{-1}$                               (d) An acid chloride,  $1780\text{ cm}^{-1}$
- 12.28** How might you use IR spectroscopy to distinguish between the three isomers 1-butyne, 1,3-butadiene, and 2-butyne?
- 12.29** Would you expect two enantiomers such as (*R*)-2-bromobutane and (*S*)-2-bromobutane to have identical or different IR spectra? Explain.
- 12.30** Would you expect two diastereomers such as *meso*-2,3-dibromobutane and (*2R,3R*)-dibromobutane to have identical or different IR spectra? Explain.
- 12.31** Propose structures for compounds that meet these descriptions:  
 (a)  $\text{C}_5\text{H}_8$ , with IR absorptions at  $3300$  and  $2150\text{ cm}^{-1}$   
 (b)  $\text{C}_4\text{H}_8\text{O}$ , with a strong IR absorption at  $3400\text{ cm}^{-1}$   
 (c)  $\text{C}_4\text{H}_8\text{O}$ , with a strong IR absorption at  $1715\text{ cm}^{-1}$   
 (d)  $\text{C}_8\text{H}_{10}$ , with IR absorptions at  $1600$  and  $1500\text{ cm}^{-1}$
- 12.32** How would you use infrared spectroscopy to distinguish between these pairs of isomers?  
 (a)  $(\text{CH}_3)_3\text{N}$  and  $\text{CH}_3\text{CH}_2\text{NHCH}_3$                       (b)  $\text{CH}_3\text{COCH}_3$  and  $\text{CH}_2=\text{CHCH}_2\text{OH}$   
 (c)  $\text{CH}_3\text{COCH}_3$  and  $\text{CH}_3\text{CH}_2\text{CHO}$
- 12.33** Two infrared spectra are shown. One is the spectrum of cyclohexane, and the other is the spectrum of cyclohexene. Identify them and explain your answer.



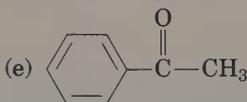
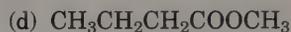
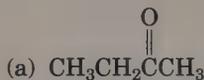
12.34 How would you use infrared spectroscopy to distinguish between these pairs of constitutional isomers?



12.35 Assume you are carrying out the dehydration of 1-methylcyclohexanol to 1-methylcyclohexene. How could you use infrared spectroscopy to determine when the reaction was complete?

12.36 Assume that you are carrying out the base-induced dehydrobromination of 3-bromo-3-methylpentane (Section 11.11). How could you use IR spectroscopy to tell which of two possible elimination products was formed?

12.37 At what approximate positions might these compounds show IR absorptions?

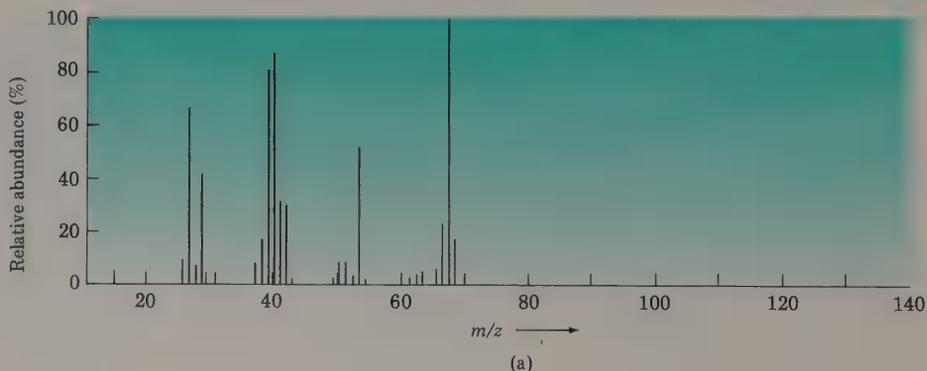


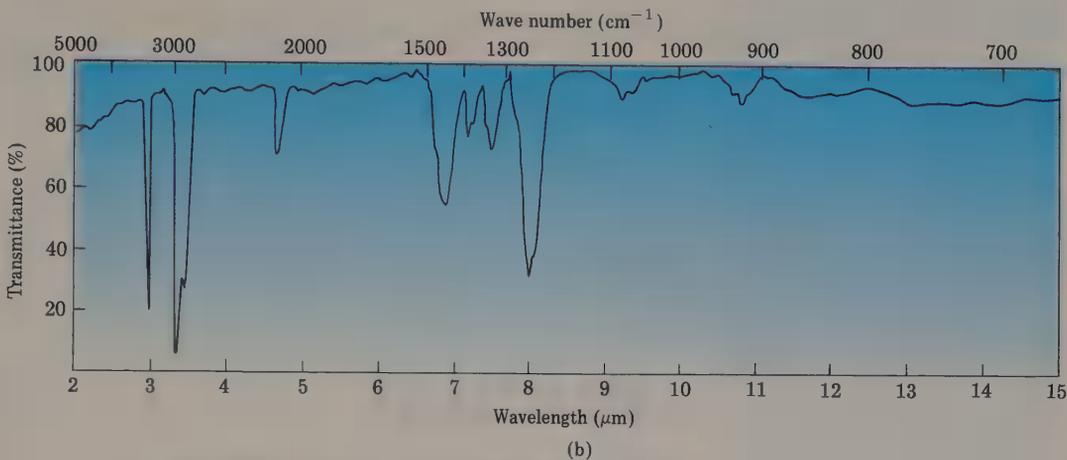
12.38 Which kind of C=O bond is stronger, that in an ester ( $1735\text{ cm}^{-1}$ ) or that in a saturated ketone ( $1715\text{ cm}^{-1}$ )? Explain your answer.

12.39 Carvone is an unsaturated ketone responsible for the odor of spearmint. If carvone has  $M^+ = 150$  in its mass spectrum, what molecular formulas are likely? If carvone has three double bonds and one ring, what molecular formula is correct?

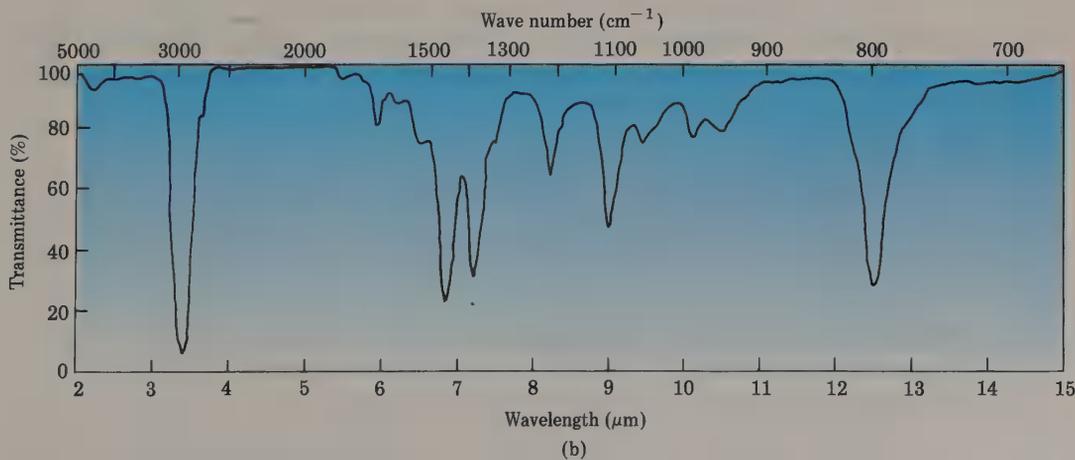
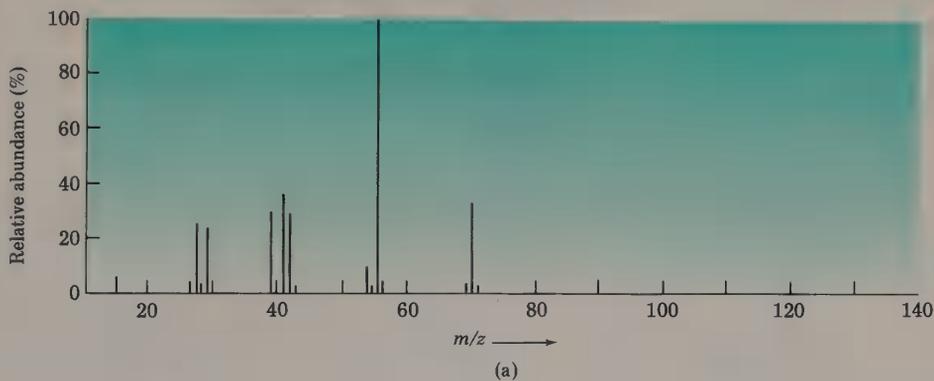
12.40 Carvone (Problem 12.39) has an intense infrared absorption at  $1690\text{ cm}^{-1}$ . What kind of ketone does carvone contain?

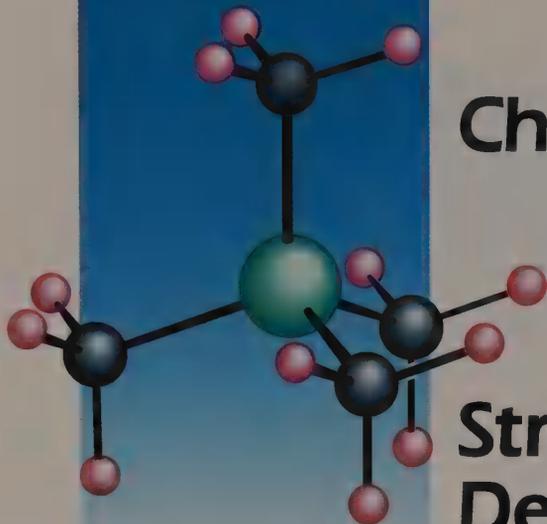
12.41 Shown are the mass spectrum (a) and the infrared spectrum (b) of an unknown hydrocarbon. Analyze the data and propose as many reasonable structures as you can.





**12.42** Shown are the mass spectrum (a) and the infrared spectrum (b) of another unknown hydrocarbon. Analyze the data and propose as many reasonable structures as you can.





## Chapter 13

# Structure Determination: Nuclear Magnetic Resonance Spectroscopy

Nuclear magnetic resonance spectroscopy (NMR) is the most valuable spectroscopic technique available to organic chemists. It's the method of structure determination that organic chemists first turn to for information.

We saw in Chapter 12 that mass spectroscopy provides information about the molecular weight and formula of a molecule and that infrared spectroscopy provides information about a molecule's functional groups. Nuclear magnetic resonance spectroscopy doesn't replace or duplicate either of these techniques; rather, it complements them by providing a "map" of the carbon-hydrogen framework of an organic molecule. Taken together, NMR, IR, and mass spectroscopy often make it possible to obtain complete solutions for the structures of complex unknowns.

Mass spectroscopy	Molecular size and formula
Infrared spectroscopy	Functional groups present
NMR spectroscopy	Map of carbon-hydrogen framework

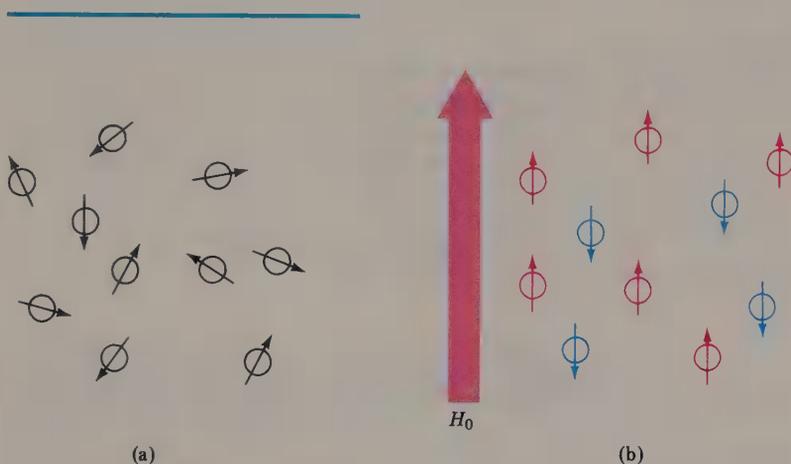
### 13.1 Nuclear Magnetic Resonance Spectroscopy

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Many kinds of nuclei behave as if they were spinning about an axis. Since they're positively charged, these spinning nuclei act like tiny magnets and therefore interact with an externally applied magnetic field (denoted  $H_0$ ).

Not all nuclei act this way, but fortunately for organic chemists, both the proton ( $^1\text{H}$ ) and the  $^{13}\text{C}$  nucleus do have spins. Let's see what the consequences of nuclear spin are and how we can use the results.

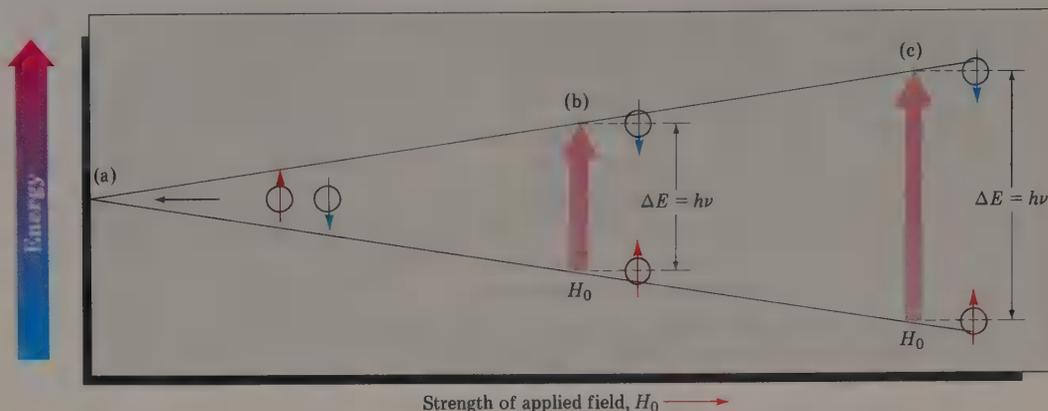
In the absence of a strong external magnetic field, the nuclear spins of magnetic nuclei are oriented randomly. When a sample containing these nuclei is placed between the poles of a strong magnet, however, the nuclei adopt specific orientations, much as a compass needle orients itself in the earth's magnetic field. A spinning  $^1\text{H}$  or  $^{13}\text{C}$  nucleus can orient so that its own tiny magnetic field is aligned either with (parallel to) or against (antiparallel to) the external field. The two orientations don't have the same energy and therefore aren't present in equal amounts. The parallel orientation is slightly lower in energy, making this spin state slightly favored over the antiparallel orientation (Figure 13.1).



**Figure 13.1** Nuclear spins are oriented randomly in the absence of a strong external magnetic field (a), but have a specific orientation in the presence of an external field,  $H_0$  (b). Note that some of the spins (red) are aligned parallel to the external field while others (blue) are antiparallel. The parallel spin state is lower in energy.

If the oriented nuclei are now irradiated with electromagnetic radiation of the proper frequency, energy absorption occurs, and the lower energy state “spin-flips” to the higher energy state. When this spin-flip occurs, the nuclei are said to be in resonance with the applied radiation, hence the name, *nuclear magnetic resonance*.

The exact amount of radio-frequency (rf) energy necessary for resonance depends both on the strength of the external magnetic field and on the identity of the nuclei being irradiated. If a very strong magnetic field is applied, the energy difference between the two spin states is large, and higher-frequency (higher-energy) radiation is required for a spin flip. If a weaker magnetic field is applied, less energy is required to effect the transition between nuclear spin states (Figure 13.2).



**Figure 13.2** The variation of energy difference between nuclear spin states as a function of applied magnetic field strength. Absorption of rf energy of frequency  $\nu$  converts a nucleus from a lower spin state to a higher spin state. Spin states (a) have equal energies in the absence of an applied magnetic field but (b) have unequal energies in the presence of a magnetic field. At  $\nu = 60$  MHz,  $\Delta E = 5.7 \times 10^{-6}$  kcal/mol. (c) The energy difference between spin states is greater at larger applied fields. At  $\nu = 100$  MHz,  $\Delta E = 9.5 \times 10^{-6}$  kcal/mol.

In practice, superconducting magnets that produce enormously powerful fields up to 140,000 gauss are sometimes used, but a field strength of 14,100 gauss is more common. At this magnetic field strength, rf energy in the 60 MHz range (1 MHz = 1 megahertz = 1 million cycles per second) is required to bring a  $^1\text{H}$  nucleus into resonance, and rf energy of 15 MHz is required to bring a  $^{13}\text{C}$  nucleus into resonance. These energy levels needed for NMR are much smaller than those required for infrared spectroscopy (1.1–11 kcal/mol); 60 MHz rf energy corresponds to only  $5.7 \times 10^{-6}$  kcal/mol ( $2.4 \times 10^{-5}$  kJ/mol).

The  $^1\text{H}$  and  $^{13}\text{C}$  nuclei are not unique in their ability to exhibit the nuclear magnetic resonance phenomenon. All nuclei with odd-numbered masses, such as  $^1\text{H}$ ,  $^{13}\text{C}$ ,  $^{19}\text{F}$ , and  $^{31}\text{P}$ , show magnetic properties. Similarly, all nuclei with even-numbered masses but odd atomic numbers show magnetic properties ( $^2\text{H}$  and  $^{14}\text{N}$ , for example). Nuclei having both even masses and even atomic numbers ( $^{12}\text{C}$ ,  $^{16}\text{O}$ ) do not give rise to magnetic phenomena (Table 13.1).

**PROBLEM** .....

- 13.1** The exact amount of energy required to spin-flip a magnetic nucleus depends not only on the strength of the external magnetic field but also on the intrinsic properties of the specific isotope. We saw earlier that, at a field strength of 14,100 gauss, rf energy of 60 MHz is required to bring a  $^1\text{H}$  nucleus into resonance. At the same field strength, rf energy of 56 MHz will bring a  $^{19}\text{F}$  nucleus into resonance. Use the

**TABLE 13.1** The NMR Behavior of Some Common Nuclei

<i>Magnetic nuclei</i>	<i>Nonmagnetic nuclei</i>
$^1\text{H}$ $^{13}\text{C}$ $^2\text{H}$ $^{14}\text{N}$ $^{19}\text{F}$ $^{31}\text{P}$	$^{12}\text{C}$ $^{16}\text{O}$ $^{32}\text{S}$
} NMR observed	} No NMR observed

equation given in Problem 12.10 on page 427 to calculate the amount of energy required to spin-flip a  $^{19}\text{F}$  nucleus. Is this amount greater or less than that required to spin-flip a  $^1\text{H}$  nucleus?

**PROBLEM** .....

- 13.2** Calculate the amount of energy required to spin-flip a proton in a spectrometer operating at 100 MHz. Does increasing the spectrometer frequency from 60 MHz to 100 MHz increase or decrease the amount of energy necessary for resonance?
- .....

## 13.2 The Nature of NMR Absorptions

From the description given thus far, you might expect all protons in a molecule to absorb rf energy at the same frequency and all  $^{13}\text{C}$  nuclei to absorb at the same frequency. If this were true, we would observe only a single NMR absorption band in the  $^1\text{H}$  or  $^{13}\text{C}$  spectrum of an unknown, a situation that would be of little use for structure determination. In fact, the absorption frequency is not the same for all nuclei.

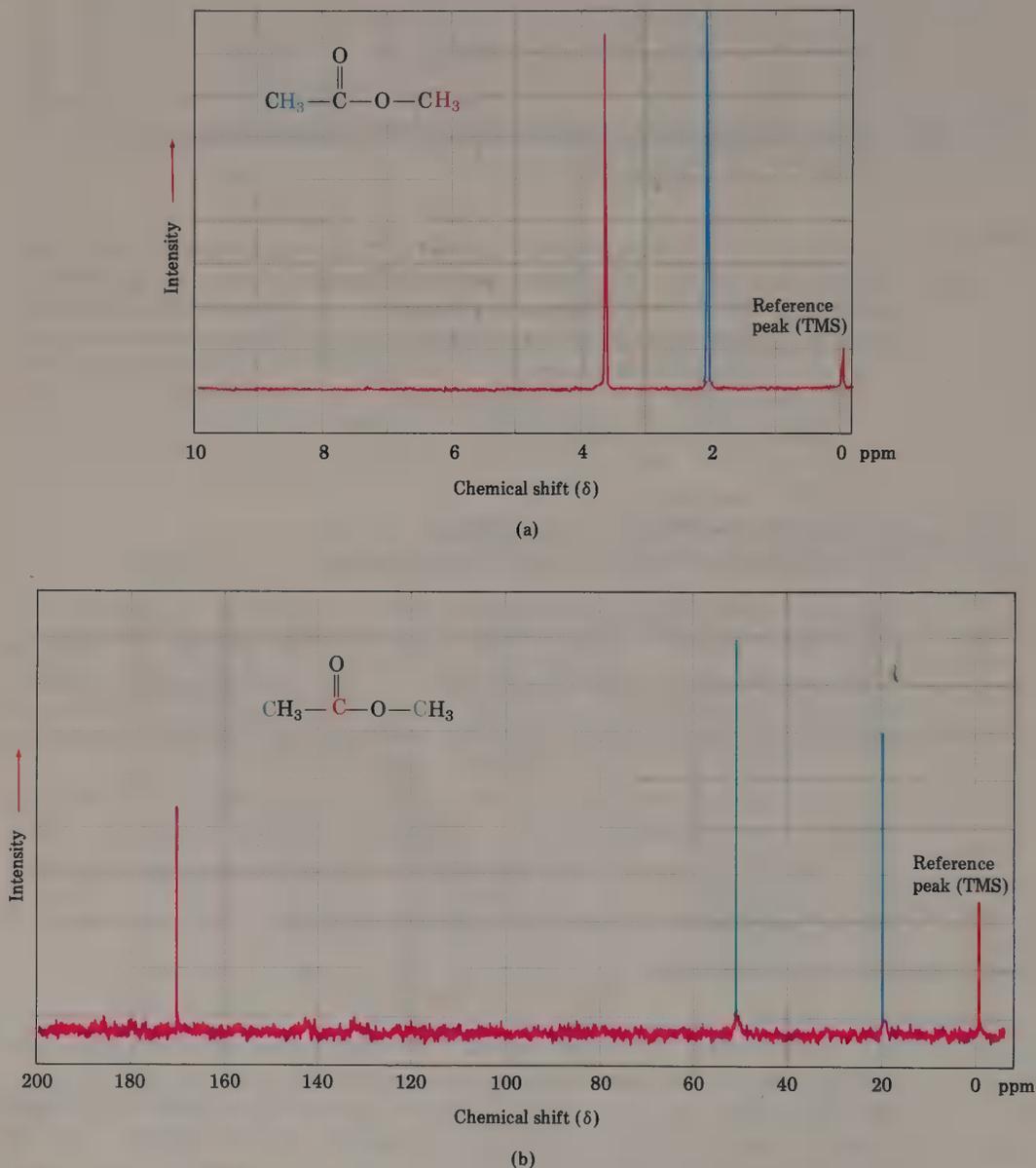
All nuclei in molecules are surrounded by electron clouds. When a uniform external magnetic field is applied to a sample molecule, the circulating electron clouds set up tiny local magnetic fields of their own. These local magnetic fields act in opposition to the applied field so that the effective field actually felt by the nucleus is a bit smaller than the applied field.

$$H_{\text{effective}} = H_{\text{applied}} - H_{\text{local}}$$

In describing this effect, we say that nuclei are **shielded** from the applied field by the circulating electron clouds that surround them. Since each kind of nucleus in a molecule is in a slightly different electronic environment, each nucleus is shielded to a slightly different extent, and the effective magnetic field actually felt is not the same for each nucleus. If the NMR instrument is sensitive enough, the tiny differences in the effective magnetic fields experienced by different nuclei can be observed, and we can see different NMR signals for each different kind of nucleus. Thus, the NMR

spectrum of an organic compound provides us with a map of the carbon-hydrogen framework. With practice, it's possible to read the map and thereby derive structural information about an unknown molecule.

Figure 13.3 shows both the  $^1\text{H}$  and the  $^{13}\text{C}$  NMR spectra of methyl acetate,  $\text{CH}_3\text{CO}_2\text{CH}_3$ . The horizontal axis shows the difference in effective field strength felt by the nuclei, and the vertical axis indicates intensity of



**Figure 13.3** (a) The  $^1\text{H}$  NMR spectrum and (b) the  $^{13}\text{C}$  NMR spectrum of methyl acetate,  $\text{CH}_3\text{CO}_2\text{CH}_3$ .

absorption of radio-frequency energy. Each peak in the NMR spectrum corresponds to a different kind of nucleus in a molecule. Note, though, that  $^1\text{H}$  and  $^{13}\text{C}$  spectra can't both be observed at the same time on the same spectrometer because different amounts of energy are required to spin-flip the different kinds of nuclei. The two kinds of spectra must be recorded separately.

The  $^{13}\text{C}$  spectrum of methyl acetate in Figure 13.3 shows three peaks, one for each of the three different carbon atoms in the molecule. The  $^1\text{H}$  NMR spectrum shows only *two* peaks, however, even though methyl acetate has six protons. One peak is due to the  $\text{CH}_3\text{CO}$  protons, and the other to the  $\text{OCH}_3$  protons. Since the three protons of each methyl group have the same chemical (and magnetic) environment, they are shielded to the same extent, are said to be *equivalent*, and show a single absorption. The two methyl groups themselves, however, are not equivalent; they therefore absorb at different positions.

The operation of an NMR spectrometer is illustrated schematically in Figure 13.4. An organic sample is dissolved in a suitable solvent and placed in a thin glass tube between the poles of a magnet. The strong magnetic field causes the  $^1\text{H}$  and  $^{13}\text{C}$  nuclei in the molecule to align in one of the two possible orientations, and the sample is irradiated with rf energy. The exact amount of energy required depends both on the strength of the magnetic field and on the kind of nucleus we intend to observe. As noted before, the two most commonly observed nuclei,  $^1\text{H}$  and  $^{13}\text{C}$ , absorb in quite different rf ranges and can't both be observed at the same time.

If the frequency of rf irradiation is held constant and the strength of the applied magnetic field is changed, each nucleus comes into resonance at a slightly different field strength. A sensitive detector monitors the absorption of rf energy, and the electronic signal is then amplified and displayed as a peak on a recorder chart.

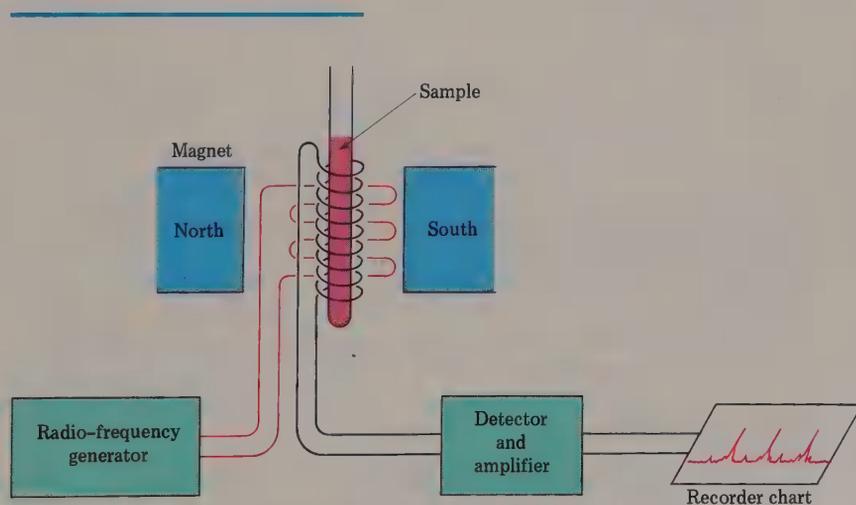


Figure 13.4 Schematic operation of an NMR spectrometer.

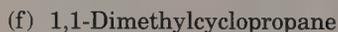
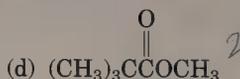
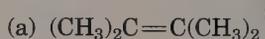
Nuclear magnetic resonance spectroscopy differs from infrared spectroscopy (Sections 12.9–12.12) in that the time scales of the two techniques are quite different. The absorption of infrared energy by a molecule giving rise to a change in vibrational states is an essentially instantaneous process (about  $10^{-13}$  sec). The NMR process, however, requires much more time (about  $10^{-3}$  sec).

The difference in time scales between IR and NMR spectroscopy can be compared to the difference between a camera operating at a very fast shutter speed and a camera operating at a very slow shutter speed. The fast camera (infrared) takes an instantaneous picture and “freezes” the action. If two rapidly interconverting species are present in a sample, infrared will record the spectrum of each. The slow camera (NMR), however, takes a blurred, “time-averaged” picture. If two species interconverting faster than  $10^3$  times per second are present in a sample, NMR will record only a single, averaged spectrum, rather than separate spectra of the two discrete species.

Because of this “blurring” effect, NMR spectroscopy can sometimes be used to measure the rate of very fast processes such as conformational changes. The NMR spectrum of the molecule undergoing the fast process is recorded at varying temperatures until the point is found where blurring of peaks occurs. By then using the known frequency of radiation, it’s possible to calculate the rate of the process under investigation.

PROBLEM .....

13.3 How many signals would you expect each of the following molecules to have in its  $^1\text{H}$  and  $^{13}\text{C}$  spectra?



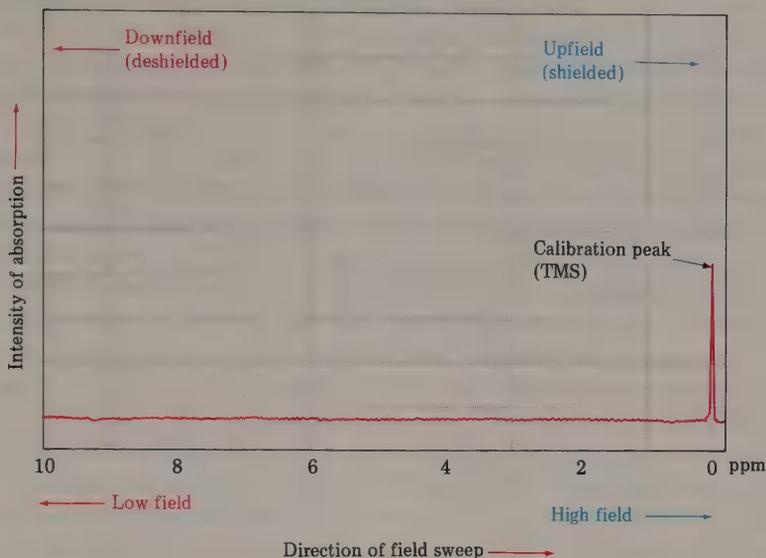
PROBLEM .....

13.4 How can you explain the fact that 2-chloropropene shows signals for three kinds of protons in its  $^1\text{H}$  NMR spectrum?

.....

## 13.3 Chemical Shifts

NMR spectra are displayed on charts that show the applied field strength increasing from left to right (Figure 13.5). Thus, the left part of the chart is the low-field (or **downfield**) side, and the right part is the high-field (or



**Figure 13.5** The NMR chart. The low-field side is on the left, and the high-field side is on the right.

**upfield**) side. In order to define the position of an absorption, the NMR chart is calibrated and a reference point is used. In practice, a small amount of tetramethylsilane [TMS,  $(\text{CH}_3)_4\text{Si}$ ] is added to the sample so that a standard reference absorption line is produced when the spectrum is run. Tetramethylsilane is used as reference for both  $^1\text{H}$  and  $^{13}\text{C}$  measurements because it gives rise in both kinds of spectra to a single peak that occurs upfield (farther right on the chart) of other absorptions normally found in organic compounds. The  $^1\text{H}$  and  $^{13}\text{C}$  spectra of methyl acetate in Figure 13.3 have the TMS reference peak indicated.

The exact place on the chart at which a nucleus absorbs is called its **chemical shift**. By convention, the chemical shift of TMS is arbitrarily set as the zero point, and other absorptions normally occur downfield (to the left on the chart). For historical reasons, NMR charts are calibrated in frequency units using an arbitrary scale called the **delta scale**. One delta unit ( $\delta$ ) is equal to 1 part per million (ppm) of the spectrometer operating frequency. For example, if we were measuring the  $^1\text{H}$  NMR spectrum of a sample using an instrument operating at 60 MHz, 1  $\delta$  would be 1 ppm of 60,000,000 Hz, or 60 Hz. Similarly, if we were measuring the spectrum using a 100 MHz instrument, then 1  $\delta = 100$  Hz. The following equation can be used for any absorption:

$$\delta = \frac{\text{Observed chemical shift (number of Hz away from TMS)}}{\text{Spectrometer frequency in MHz}}$$

Although this method of calibrating NMR charts may seem needlessly complex, there's a good reason for it. As we saw earlier, the chemical shift

of an NMR peak depends on magnetic field strength, which in turn depends on the spectrometer frequency. Since many different kinds of spectrometers operating at many different magnetic field strengths and rf frequencies are available, chemical shifts given in Hz vary greatly from one instrument to another. By using a system of measurement in which NMR absorptions are expressed in *relative* terms (ppm of spectrometer frequency) rather than in absolute terms (Hz), comparisons of spectra obtained on different instruments are possible. *The chemical shift of an NMR absorption given in ppm or  $\delta$  units is constant, regardless of the operating frequency of the spectrometer.* A  $^1\text{H}$  nucleus that absorbs at 2.0  $\delta$  on a 60 MHz instrument ( $2.0 \text{ ppm} \times 60 \text{ MHz} = 120 \text{ Hz}$  downfield from TMS) also absorbs at 2.0  $\delta$  on a 300 MHz instrument ( $2.0 \text{ ppm} \times 300 \text{ MHz} = 600 \text{ Hz}$  downfield from TMS).

The range in which most NMR absorptions occur is quite narrow. Almost all  $^1\text{H}$  NMR absorptions occur 0–10  $\delta$  downfield from the proton absorption of TMS, and almost all  $^{13}\text{C}$  absorptions occur 1–250  $\delta$  downfield from the carbon absorption of TMS. Thus, there is a considerable likelihood that accidental overlap of nonequivalent signals will occur. The advantage of using an instrument with high field strength (say, 300 MHz) rather than low field strength (60 MHz) is that different NMR absorptions are more widely separated at high field strength. The chances that two signals will accidentally overlap are also lessened, and interpretation of spectra becomes easier. For example, two signals that are only 6 Hz apart at 60 MHz (0.1 ppm) are 30 Hz apart at 300 MHz (still 0.1 ppm).

Let's now look more closely at the interpretation of NMR spectra to see how this tool can be used in organic structure determination. We'll begin by looking at  $^{13}\text{C}$  NMR. For technical reasons,  $^{13}\text{C}$  spectra are more difficult to obtain than  $^1\text{H}$  spectra, but they're also much easier to interpret. Although the complexities of spectrometer operation differ, the principles behind  $^{13}\text{C}$  and  $^1\text{H}$  NMR are the same. What we learn now about interpreting  $^{13}\text{C}$  spectra will simplify the subsequent discussion of  $^1\text{H}$  spectra.

PROBLEM .....

- 13.5 When the  $^1\text{H}$  NMR spectrum of acetone,  $\text{CH}_3\text{COCH}_3$ , is recorded on an instrument operating at 60 MHz, a single sharp resonance line at 2.1  $\delta$  is observed.
- How many hertz downfield from TMS does the acetone resonance line correspond to?
  - If the  $^1\text{H}$  NMR spectrum of acetone were recorded at 100 MHz, what would be the position of the absorption in  $\delta$  units?
  - How many hertz downfield from TMS does this 100 MHz spectrum correspond to?

PROBLEM .....

- 13.6 The following  $^1\text{H}$  NMR resonances were recorded on a spectrometer operating at 60 MHz. Convert each into  $\delta$  units.
- |                                     |                                       |
|-------------------------------------|---------------------------------------|
| (a) $\text{CHCl}_3$ ; 436 Hz        | (b) $\text{CH}_3\text{Cl}$ ; 183 Hz   |
| (c) $\text{CH}_3\text{OH}$ ; 208 Hz | (d) $\text{CH}_2\text{Cl}_2$ ; 318 Hz |
- .....

## 13.4 $^{13}\text{C}$ NMR Spectroscopy

In some ways, it's surprising that carbon NMR is even possible. After all,  $^{12}\text{C}$ , the most abundant carbon isotope, has no nuclear spin and isn't observable by NMR.  $^{13}\text{C}$  is the only naturally occurring carbon isotope with a spin, but its natural abundance is only about 1.1%. Thus, only about 1 of every 100 carbons in organic molecules is observable by NMR. Fortunately, the technical problems caused by this low abundance of  $^{13}\text{C}$  have been overcome through the use of improved electronics and computer techniques, and  $^{13}\text{C}$  NMR is now a routine structural tool. A  $^{13}\text{C}$  NMR spectrum can often be obtained on 10 mg of sample in a few hours.

At its simplest,  $^{13}\text{C}$  NMR makes it possible to count the different kinds of carbons in a molecule of unknown structure. In addition, we can get information about the chemical (magnetic) environment of each kind of carbon by observing its chemical shift.

Several different modes of operation are possible with  $^{13}\text{C}$  NMR instruments. In the normal mode, called the **proton noise-decoupled** mode, a single sharp resonance line is observed for each nonequivalent kind of carbon atom in a molecule, as illustrated by the spectrum of methyl acetate in Figure 13.3. There are three carbon atoms in methyl acetate and three peaks in its  $^{13}\text{C}$  NMR spectrum.

Most  $^{13}\text{C}$  resonances are between 0 and 220 ppm downfield from the TMS reference line, with the exact chemical shift of each  $^{13}\text{C}$  resonance dependent on that carbon's environment within the molecule. Table 13.2 and Figure 13.6 show how environment and chemical shift are correlated.

**TABLE 13.2** Carbon-13 NMR Chemical Shift Correlations

Type of carbon	Chemical shift ( $\delta$ )	Type of carbon	Chemical shift ( $\delta$ )
C—I	0–40	=C	100–150
C—Br	25–65	C—O	40–80
C—Cl	35–80	C=O	170–210
—CH <sub>3</sub>	8–30		110–160
—CH <sub>2</sub> —	15–55	C—N	30–65
	20–60		
≡C	65–85		

Although the factors that determine chemical shifts are complex, it's possible to make some generalizations from the information in Table 13.2 and Figure 13.6. One clear trend is that a carbon's chemical shift is affected

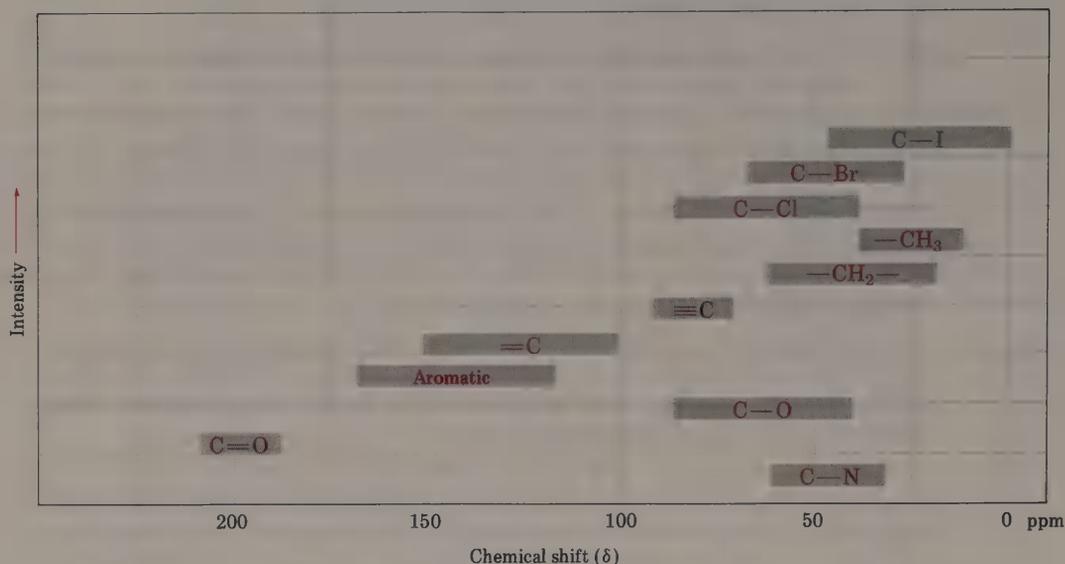


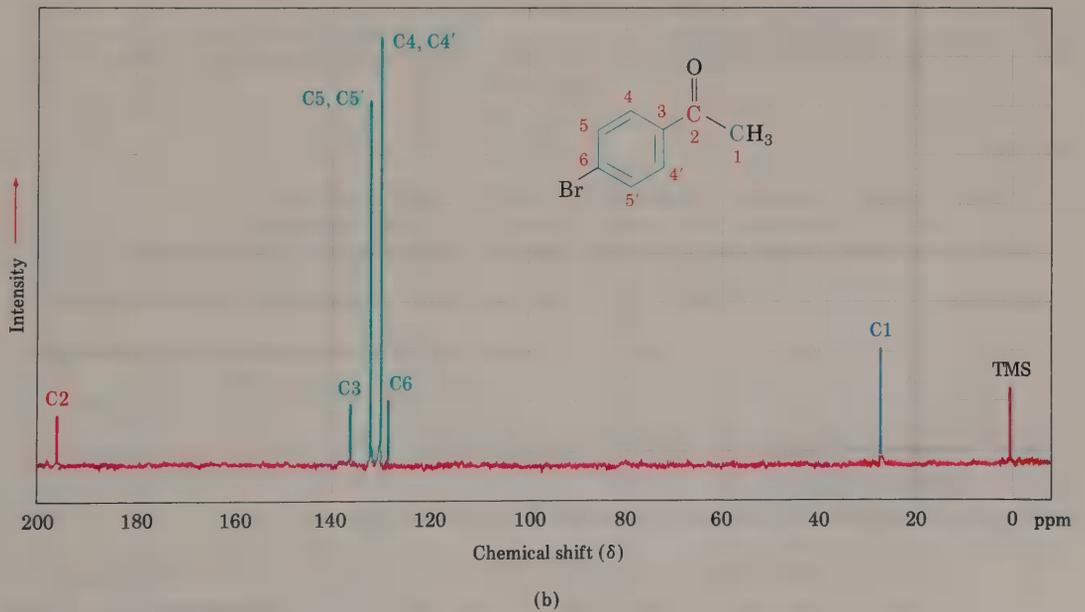
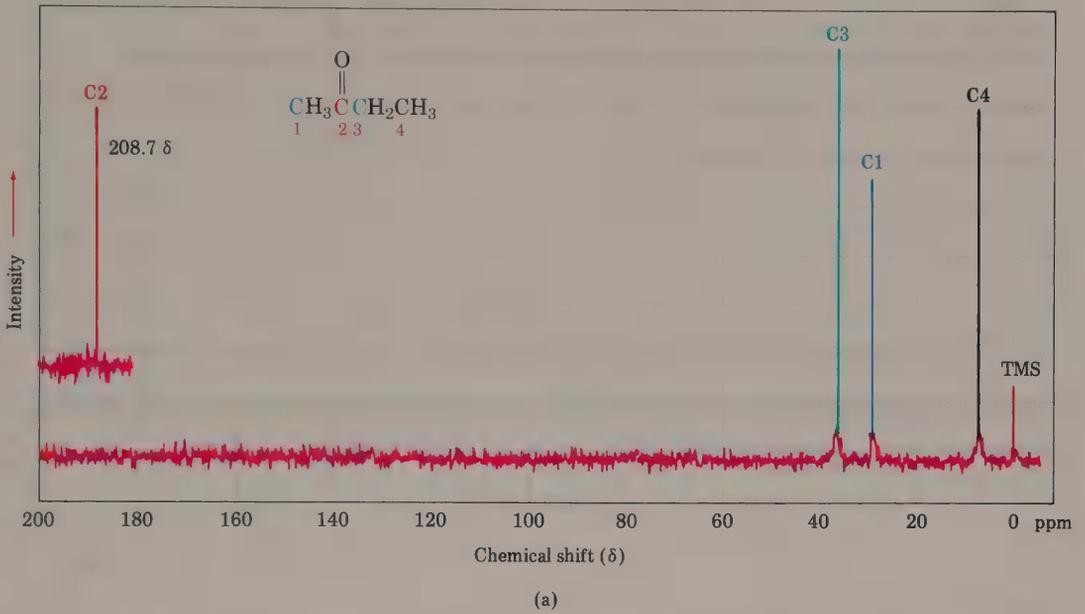
Figure 13.6 Chemical shift correlations for  $^{13}\text{C}$  NMR.

by the electronegativity of the atoms it's bonded to. Carbons bonded to electronegative atoms like oxygen, nitrogen, or halogen absorb downfield (to the left) of normal alkane carbons. This trend is easy to explain if you think for a moment about the causes of chemical shifts. Since electronegative atoms attract electrons, they pull electrons away from neighboring carbons, causing the carbons to be deshielded and to come into resonance at a lower field than alkane carbons.

Another, less easily explained trend is that  $sp^3$ -hybridized carbons generally absorb in the 0–90  $\delta$  range and  $sp^2$  carbons absorb in the 100–210  $\delta$  range. Carbonyl carbons (C=O) are particularly distinct in  $^{13}\text{C}$  NMR and are easily observed at the extreme low-field end of the spectrum in the range 170–210  $\delta$ . Figure 13.7 shows the  $^{13}\text{C}$  NMR spectra of 2-butanone and *para*-bromoacetophenone, and indicates the peak assignments. Note that the carbonyl carbons are at the left edge of the spectrum in each case.

The  $^{13}\text{C}$  NMR spectrum of *para*-bromoacetophenone is interesting in several ways. Note particularly that only six carbon absorptions are observed even though the molecule contains eight carbons. *para*-Bromoacetophenone has a symmetry plane that makes ring carbons 4 and 4', and carbons 5 and 5', equivalent. Thus, the six ring carbons show only four absorptions in the range 128–137  $\delta$ .

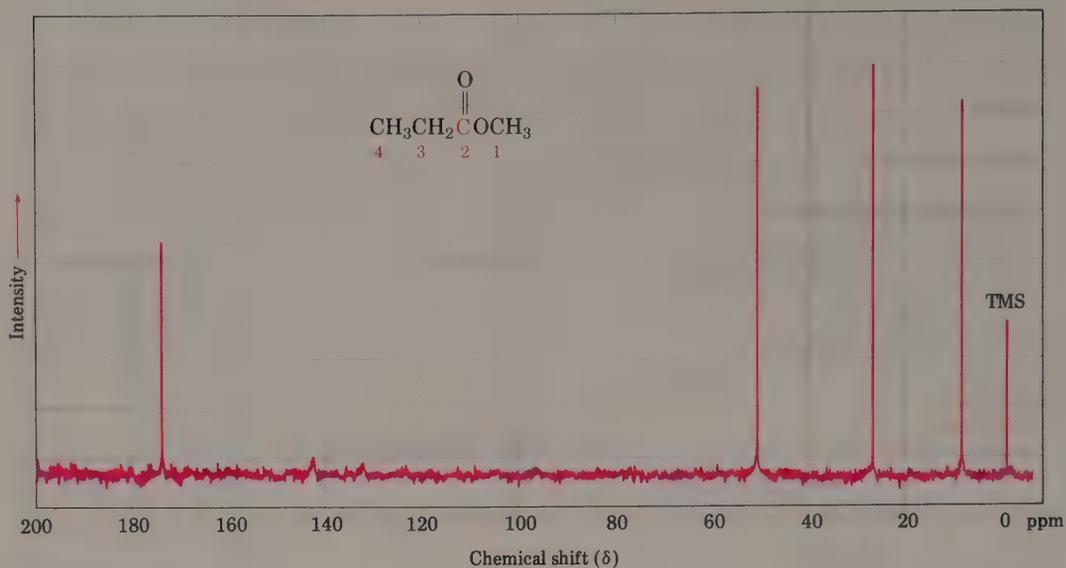
Another interesting feature of both spectra is that the peaks aren't uniform in size. Some peaks are larger than others even though all are one-carbon resonances (except for the 2 two-carbon peaks of *para*-bromoacetophenone). This difference in peak size is caused by complex factors that we won't go into and is only indirectly related to the structure of the sample.



**Figure 13.7** Carbon-13 NMR spectra of (a) 2-butanone and (b) *para*-bromoacetophenone.

**PROBLEM** .....

**13.7** Assign the resonances in the following  $^{13}\text{C}$  NMR spectrum of methyl propanoate,  $\text{CH}_3\text{CH}_2\text{CO}_2\text{CH}_3$ .



PROBLEM .....

13.8 Predict the number of carbon resonance lines you would expect to observe in the  $^{13}\text{C}$  NMR spectra of these compounds:

- |                         |                         |
|-------------------------|-------------------------|
| (a) Methylcyclopentane  | (b) 1-Methylcyclohexene |
| (c) 1,2-Dimethylbenzene | (d) 2-Methyl-2-butene   |

PROBLEM .....

13.9 Propose structures for compounds that fit these descriptions:

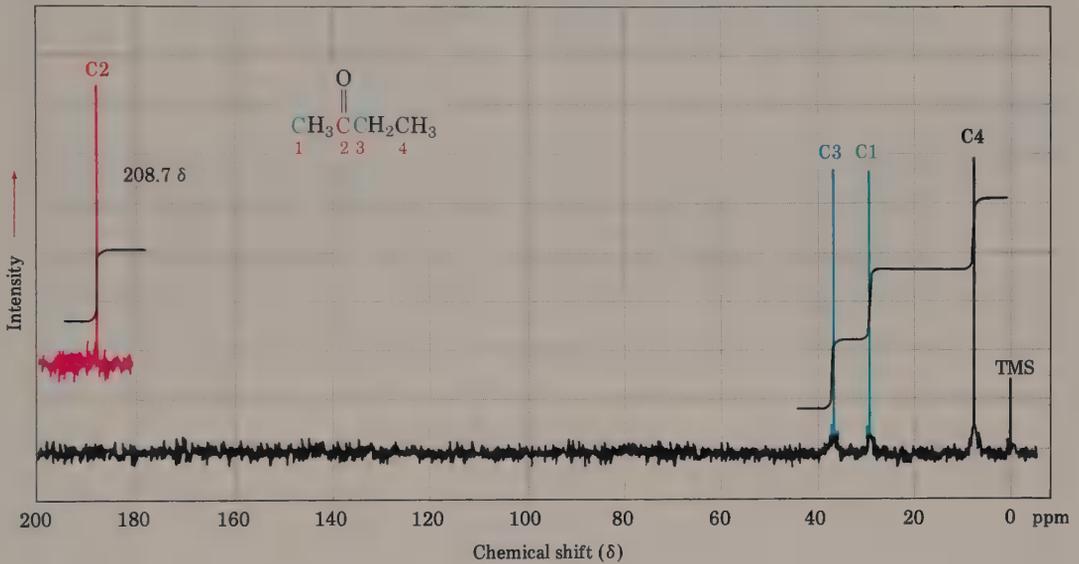
- A hydrocarbon with seven lines in its  $^{13}\text{C}$  NMR spectrum
  - A six-carbon compound with only five lines in its  $^{13}\text{C}$  NMR spectrum
  - A four-carbon compound with three lines in its  $^{13}\text{C}$  NMR spectrum
- .....

## 13.5 Measurement of NMR Peak Areas: Integration

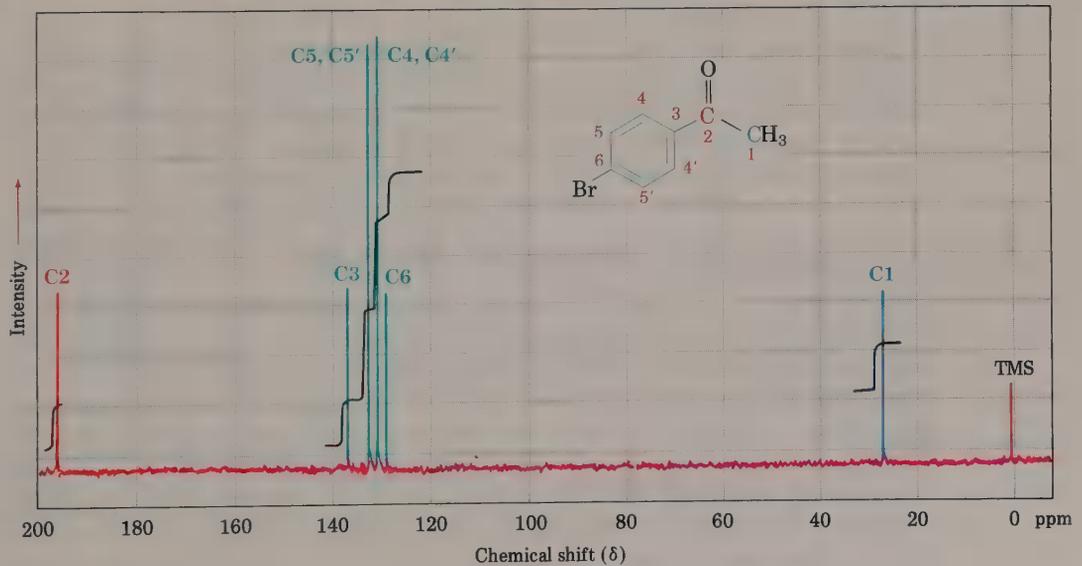
The relative sizes of the different peaks observed in a  $^{13}\text{C}$  NMR spectrum depend to a considerable extent on the mode of spectrometer operation. When the spectrometer is operated in the normal mode, different single-carbon peaks often have different areas. When the NMR spectrometer is operated in the *integrating* mode (called the **gated-decoupled** mode), however, all single-carbon resonances have the same peak areas. By electronically measuring, or integrating, the area under each peak, it's possible to tell the relative number of carbon atoms each peak represents.

Integrated peak areas are presented on the chart in a stair-step manner, with the height of each step proportional to the number of carbons causing that peak. To compare one peak size against another, you simply take out

a ruler and measure the heights of the various steps. Figure 13.8 shows the integrated  $^{13}\text{C}$  NMR spectra of 2-butanone and *para*-bromoacetophenone. Note that all four peaks in the 2-butanone spectrum have the same integrated peak area, while the *para*-bromoacetophenone spectrum shows 4 equal one-carbon resonance lines and 2 two-carbon resonance lines.



(a)



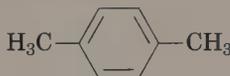
(b)

**Figure 13.8** Integrated  $^{13}\text{C}$  NMR spectra of (a) 2-butanone and (b) *para*-bromoacetophenone.

Integrated spectra give more information than normal spectra do, but this information comes at a price. The NMR spectrometer is two or three times less sensitive in the integrating mode than in the normal mode so that more sample and more time are required to obtain the spectrum. Since this price is not worth paying unless a particular ambiguity exists about the number of carbon atoms in a sample, integrated  $^{13}\text{C}$  spectra are not often obtained in practice. The concept of NMR spectra integration is a useful one, though, and we'll soon see that peak integration is critically important in interpreting  $^1\text{H}$  NMR spectra.

PROBLEM .....

- 13.10 How many lines would you expect to see in the  $^{13}\text{C}$  NMR spectrum of *para*-dimethylbenzene? What peak areas would you expect to find on integration of the spectrum? Refer to Table 13.2 for approximate values of chemical shifts, and then sketch what the integrated spectrum might look like.



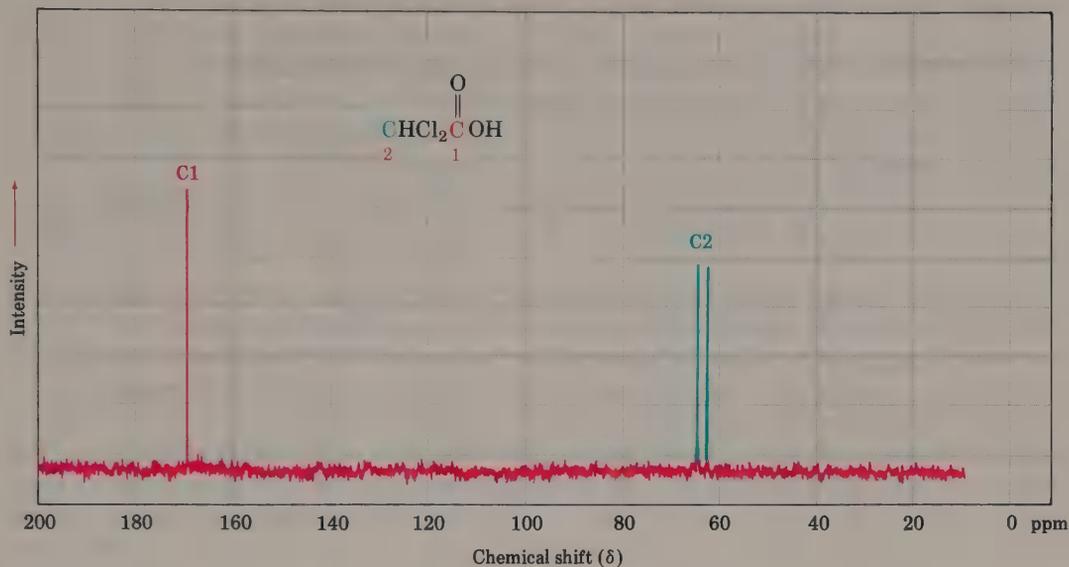
*para*-Dimethylbenzene

## 13.6 Spin–Spin Splitting of NMR Signals

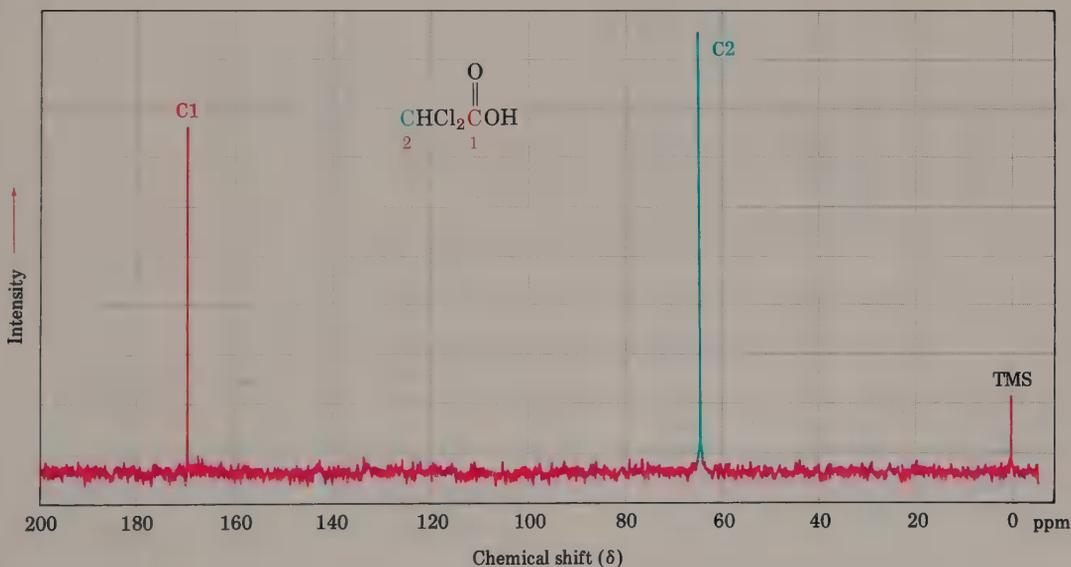
For most purposes, the normal  $^{13}\text{C}$  NMR spectrum is fully satisfactory. Obtaining the spectrum in this mode provides a carbon count of the sample molecule and gives information about the environment of each carbon. Occasionally, however, more detailed information is needed, and yet a third mode of spectrometer operation is employed. When the spectrometer is operated in the *spin-coupled* mode (called the **off-resonance** mode), single-carbon resonance lines can split into multiple lines. What is this signal splitting due to, and what use can we make of this phenomenon?

Look carefully at both the normal spectrum and the spin-coupled spectrum of dichloroacetic acid shown in Figure 13.9. When the two spectra are displayed on the same chart, it becomes clear that the carbonyl-carbon resonance at  $170\ \delta$  remains a *singlet*, but the methine-carbon resonance ( $\text{R}_3\text{C-H}$ ) at  $64\ \delta$  splits into two peaks (a *doublet*). This phenomenon, known as **spin–spin splitting**, is due to the fact that the nuclear spin of one atom interacts, or **couples**, with the nuclear spin of a nearby atom. In other words, the tiny magnetic field of one nucleus affects the magnetic field felt by a neighboring nucleus.

How does spin–spin splitting arise? Let's review briefly. We've seen that, when a magnetic nucleus such as  $^{13}\text{C}$  is placed in a strong magnetic field, the nucleus adopts one of two spin states. The  $^{13}\text{C}$  spin lines up either with or against the applied magnetic field. In addition to orienting the nuclear



(a)



(b)

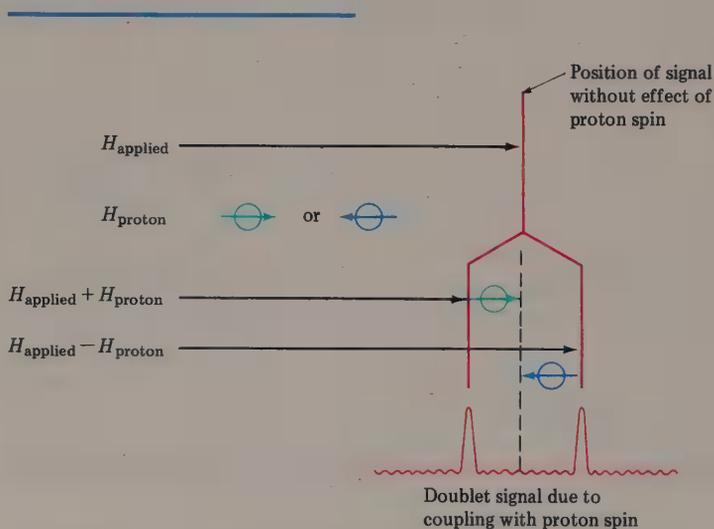
**Figure 13.9** Normal spectrum (lower scan) and spin-coupled spectrum (upper scan) of dichloroacetic acid,  $\text{CHCl}_2\text{COOH}$ .

spins, the applied magnetic field also causes electrons in the molecule to shield the nucleus by setting up tiny local magnetic fields that act in opposition to the applied field. Because of this shielding, the effective field actually felt by the nucleus is slightly less than the applied field. Differences in

the extent of shielding at each nucleus account for the differences in chemical shifts between nuclei that we observe in  $^{13}\text{C}$  NMR spectra.

In addition to being affected by electron shielding, the magnetic field felt by a nucleus is also affected by neighboring magnetic nuclei. For example, the C=O carbon of dichloroacetic acid is bonded only to nonmagnetic neighbor atoms, but the  $-\text{CHCl}_2$  carbon is bonded to another magnetic nucleus—a proton,  $^1\text{H}$ . (Remember that the natural abundance of  $^{13}\text{C}$  is only about 1.1%, and the likelihood of two  $^{13}\text{C}$  atoms being adjacent in a molecule is therefore small. Thus, we don't observe  $^{13}\text{C}$ – $^{13}\text{C}$  magnetic interactions.)

When placed in a strong external magnetic field, the spin of the neighboring  $^1\text{H}$  nucleus aligns either with or against the applied field. If the neighboring spin is aligned *with* the applied field, the total effective field at the neighboring carbon is slightly larger than it would otherwise be. Consequently, the applied field necessary to cause resonance is slightly reduced. Conversely, if the neighboring spin is aligned *against* the applied field, the effective field at the neighboring carbon is slightly smaller than it would otherwise be. Thus, the applied field needed to bring the carbon into resonance is slightly increased. Figure 13.10 shows schematically how spin–spin splitting arises.



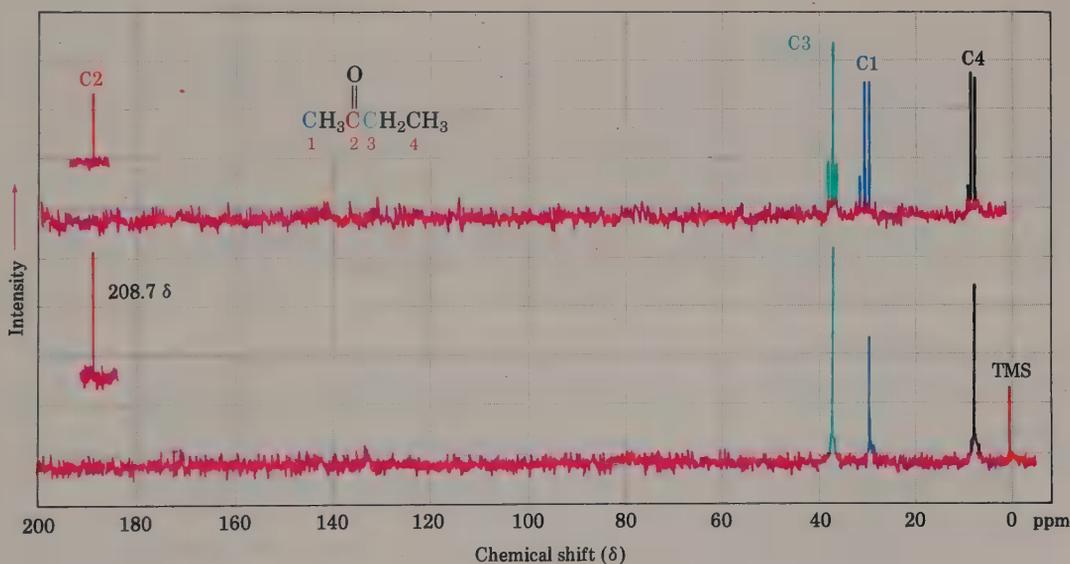
**Figure 13.10** Origin of spin–spin splitting. If the neighboring proton has its spin aligned *with* the applied field, then a lower field value is required for resonance. If the neighboring proton has its spin *against* the applied field, then a higher value is required. Resonance therefore occurs at two different values of the applied field, and a doublet absorption is observed.

The consequence of the coupling of adjacent nuclear spins is that the carbon comes into resonance at two slightly different values of the applied field. One resonance is a little above where it would be without coupling,

and the other resonance is a little below where it would be without coupling. We therefore observe a doublet in the  $^{13}\text{C}$  NMR spectrum of dichloroacetic acid.

It's important to realize that the signal from an *individual* carbon atom in an individual molecule is not a doublet. An individual carbon atom absorbs at either one position or the other depending on the orientation of its own adjacent hydrogen. We see a doublet in the NMR spectrum only because we are observing the summation of signals from a great many molecules.

If a carbon atom is bonded to more than one proton, more complex spin-spin splitting is observed. For example, Figure 13.11 shows both the normal spectrum and the spin-coupled spectrum of 2-butanone. The carbonyl-carbon (C2) resonance at  $208\ \delta$  because it's not bonded to any protons. The  $-\text{CH}_2-$  resonance at  $36.8\ \delta$  splits into a *triplet*, however, and the two  $-\text{CH}_3$  resonances at  $6.6\ \delta$  and at  $28.8\ \delta$  become *quartets*. According to the  $n + 1$  rule, a carbon bonded to  $n$  protons shows  $n + 1$  peaks in its spin-coupled  $^{13}\text{C}$  spectrum. How does this rule arise?

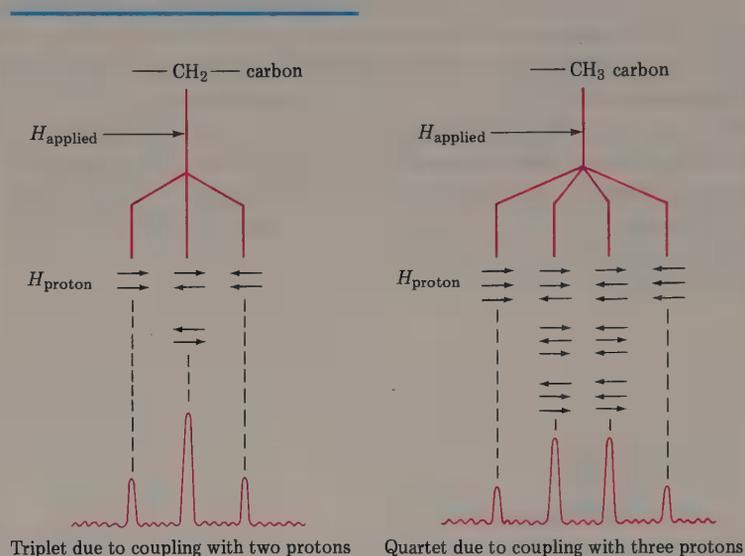


**Figure 13.11** Normal spectrum (lower scan) and spin-coupled spectrum (upper scan) of 2-butanone.

A methylene carbon ( $-\text{CH}_2-$ ) is bonded to two other magnetic nuclei (protons), and the spins of these two protons can orient in three different combinations. Both proton spins can orient against the applied field; one can orient against and one with the applied field (two possibilities); or both

can orient with the applied field. This leads to three peaks of relative intensity 1:2:1 in the spin-coupled spectrum of a  $-\text{CH}_2-$  carbon (Figure 13.12).

A similar analysis for a  $-\text{CH}_3$  carbon predicts that spin-spin splitting should lead to the observation of a quartet of 1:3:3:1 intensity in the spin-coupled  $^{13}\text{C}$  NMR spectrum. This is, in fact, observed (Figure 13.12). To repeat: A carbon bonded to  $n$  protons gives a signal that is split into  $n + 1$  peaks in the spin-coupled spectrum.



**Figure 13.12** Spin-spin splitting of a methylene carbon ( $-\text{CH}_2-$ ) and a methyl carbon ( $-\text{CH}_3$ ). The signal of a carbon bonded to  $n$  hydrogens splits into  $n + 1$  peaks (the  $n + 1$  rule).

**PROBLEM** .....

- 13.11** Sketch what the normal  $^{13}\text{C}$  NMR spectrum of propene might look like.
- How many carbon resonances are there?
  - Into how many peaks does each carbon resonance split in the spin-coupled spectrum?

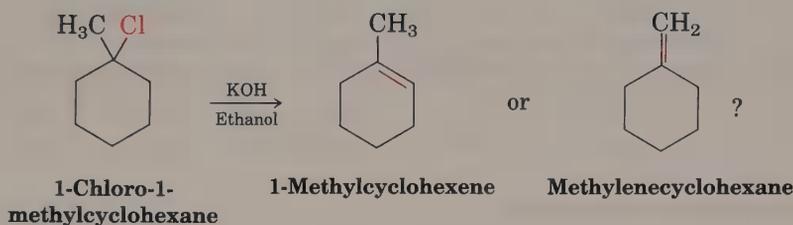
**PROBLEM** .....

- 13.12** Sketch what the normal  $^{13}\text{C}$  NMR spectrum of cyclopentene might look like.
- How many carbon resonances are there?
  - What further information would the integrated spectrum give you? Make a sketch of what the spin-coupled spectrum of cyclopentene might look like.
- .....

## 13.7 Summary of $^{13}\text{C}$ NMR Spectroscopy

The information derived from  $^{13}\text{C}$  NMR spectroscopy is extraordinarily useful for structure determination. Not only can we count the number of different kinds of carbon atoms and deduce information about their chemical environments, we can also find how many protons are bonded to each carbon. This allows us to answer many structural questions that can't be easily handled by infrared spectroscopy or mass spectroscopy.

Let's take an example. How might we prove that E2 elimination of an alkyl halide gives the more highly substituted alkene (Zaitsev's rule, Section 11.10)? Does reaction of 1-chloro-1-methylcyclohexane with strong base lead predominantly to 1-methylcyclohexene or to methylenecyclohexane?

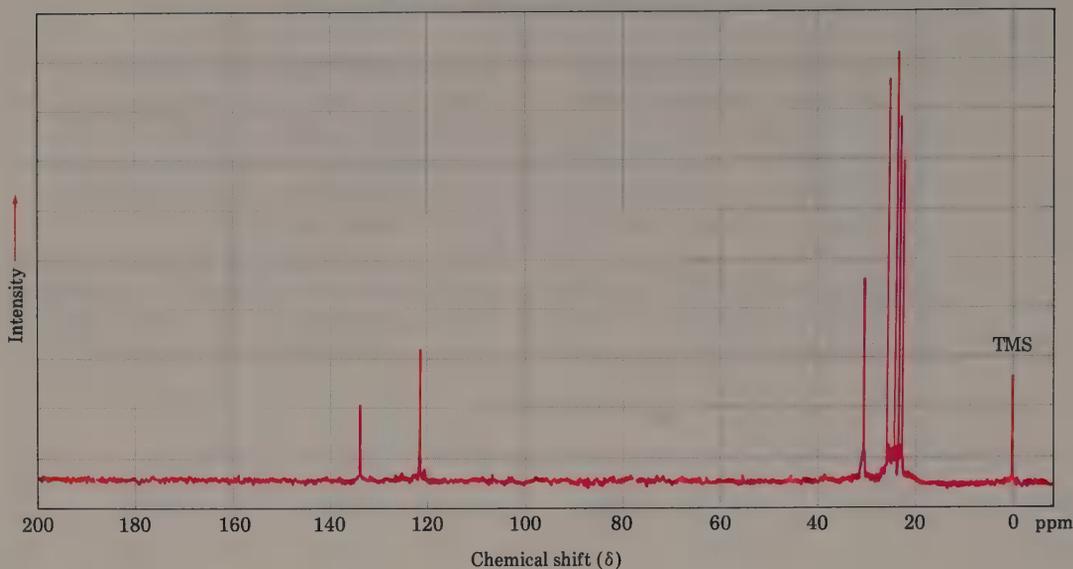


1-Methylcyclohexene should have five  $sp^3$ -carbon resonances in the range 20–50  $\delta$  and two  $sp^2$ -carbon resonances in the range 100–150  $\delta$ . Methylenecyclohexane, however, because of its symmetry, would be expected to have only three  $sp^3$ -carbon resonance peaks and two  $sp^2$ -carbon peaks. The spectrum of the actual reaction product shown in Figure 13.13 clearly identifies 1-methylcyclohexene as the substance formed in this E2 reaction.

A summary of the operating modes of  $^{13}\text{C}$  NMR spectrometers and of the kinds of information obtainable in each mode is given in Table 13.3.

**TABLE 13.3** Summary of  $^{13}\text{C}$  NMR and the Modes of Spectrometer Operation

<i>Spectrometer operating mode</i>	<i>Structural information obtained</i>
Normal mode (proton noise-decoupled)	Gives one resonance line for each kind of non-equivalent carbon atom, making it possible to count the number of kinds of carbons and deduce their environments from chemical shifts.
Integrating mode (gated-decoupled)	All resonances are of equal peak area, making it possible to integrate the area under each peak and determine the relative number of carbons each represents.
Spin-coupled mode (off-resonance)	Spin-spin splitting causes carbon resonances to split into multiplets, making it possible to determine how many protons are bonded to each carbon (the $n + 1$ rule).



**Figure 13.13** Normal  $^{13}\text{C}$  NMR spectrum of the elimination product from 1-chloro-1-methylcyclohexane.

**PROBLEM** .....

- 13.13** We saw in Section 7.6 that addition of HBr to alkenes under radical conditions leads to the non-Markovnikov product with the bromine bonding to the less highly substituted carbon. How could you use  $^{13}\text{C}$  NMR to identify the product of radical addition of HBr to 2-methylpropene?
- .....

## 13.8 $^1\text{H}$ NMR Spectroscopy

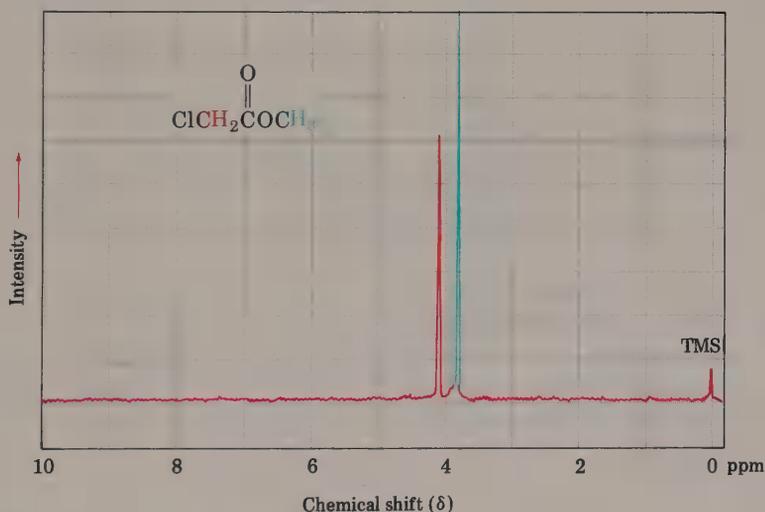
We have seen four general features of NMR spectra in the previous pages:

1. *Number of NMR absorptions* Each nonequivalent  $^{13}\text{C}$  or  $^1\text{H}$  nucleus can give rise to a separate absorption peak.
2. *Chemical shifts* The exact position of an NMR absorption tells the chemical (magnetic) environment of a given nucleus.
3. *Integration of NMR absorptions* Electronic integration of the area under a peak tells how many nuclei cause that specific resonance peak.
4. *Spin-spin splitting* The splitting of resonance lines into multiplets by coupling of nuclear spins provides information about neighboring magnetic nuclei.

All four of these features are just as applicable to  $^1\text{H}$  NMR spectra as they are to  $^{13}\text{C}$  NMR spectra. In fact, they're even easier to observe for  $^1\text{H}$  spectra because normal spectrometer operating conditions enable signal integration and spin-spin splitting. Proton NMR spectra are so easy to obtain that it usually takes only a few minutes and several milligrams of sample. Let's look at each of the four features in more detail.

## 13.9 Number of NMR Absorptions: Proton Equivalence

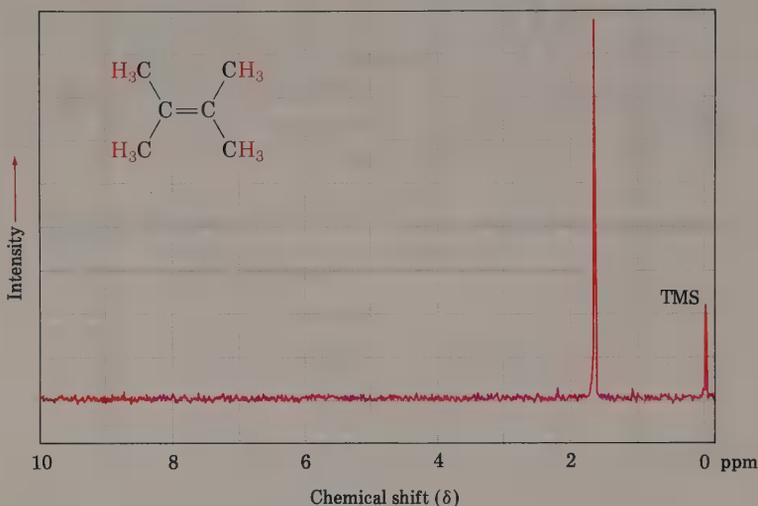
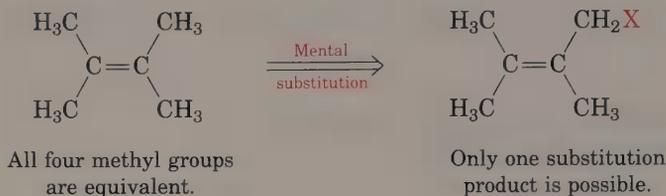
In  $^{13}\text{C}$  NMR, each unique kind of carbon atom normally gives rise to a distinct peak in the spectrum. Exactly the same is true in  $^1\text{H}$  NMR. Each unique kind of proton gives rise to its own absorption peak, and we can use this information to determine how many different kinds of protons are present. For example, Figure 13.14 shows the  $^1\text{H}$  NMR spectrum of methyl chloroacetate,  $\text{ClCH}_2\text{COCH}_3$ . There are two kinds of protons present,  $\text{Cl}-\text{CH}_2-$  protons and  $-\text{CH}_3$  protons, each of which has its own signal.



**Figure 13.14** The  $^1\text{H}$  NMR spectrum of methyl chloroacetate,  $\text{ClCH}_2\text{COCH}_3$ .

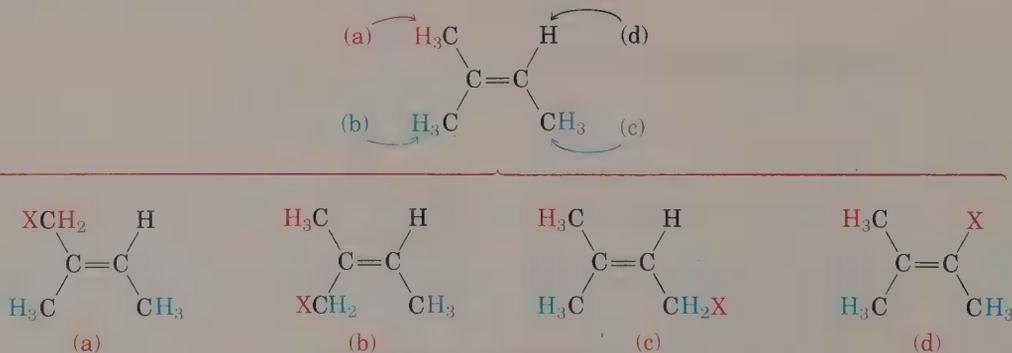
Simple visual inspection of a structure is usually enough to decide how many kinds of nonequivalent protons are present. If doubt exists, however, the equivalence or nonequivalence of protons can be determined by asking whether the same structure or different structures would be formed by mentally substituting an X group for one of the protons. If the protons are chemically equivalent, the same product would be formed no matter which proton was substituted for. If they're not chemically equivalent, different products would be formed on substitution.

For example, all 12 protons in 2,3-dimethyl-2-butene are equivalent. No matter which proton we mentally replace by an X group, we get the same structure. The 12 protons thus give rise to a single, sharp  $^1\text{H}$  NMR peak (Figure 13.15).



**Figure 13.15** The  $^1\text{H}$  NMR spectrum of 2,3-dimethyl-2-butene.

In contrast, the 10 protons of 2-methyl-2-butene are *not* all equivalent. There are three different kinds of methyl-group protons and one vinylic proton, leading to four different signals.



Four different substitution products

## PROBLEM.....

- 13.14 How many nonequivalent kinds of protons are present in these compounds?
- (a)  $\text{CH}_3\text{CH}_2\text{Br}$  (b)  $\text{CH}_3\text{OCH}_2\text{CH}(\text{CH}_3)_2$   
 (c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{NO}_2$  (d) Methylbenzene  
 (e) 2-Methyl-1-butene (f) *cis*-3-Hexene

## 13.10 Chemical Shifts in $^1\text{H}$ NMR Spectroscopy

We saw in looking at  $^{13}\text{C}$  NMR spectroscopy that differences in chemical shift are caused by the small local magnetic fields of the electrons that surround the various nuclei. Nuclei that are strongly shielded by electrons require a higher applied field to bring them into resonance (to the right on the NMR chart), and nuclei that are not strongly shielded need a lower applied field for resonance to occur (to the left on the NMR chart). Everything we've said about  $^{13}\text{C}$  chemical shifts is also applicable to  $^1\text{H}$  shifts. Proton chemical shifts tell a great deal about the chemical (magnetic) environments within a molecule.

Proton chemical shifts are expressed in  $\delta$  units ( $1 \delta = 1 \text{ ppm}$  of spectrometer frequency), just as for  $^{13}\text{C}$  shifts. In contrast to  $^{13}\text{C}$  shifts, however, proton chemical shifts fall within the rather narrow range of 0–10  $\delta$ . The exact spot within this range is highly characteristic of environment, as shown in Table 13.4 (page 466) and Figure 13.16.

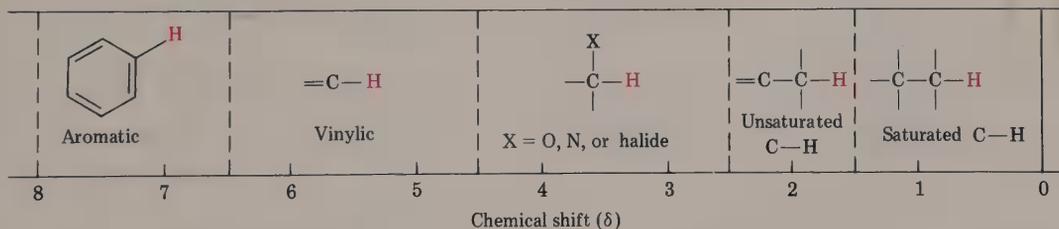


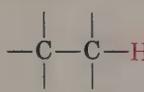
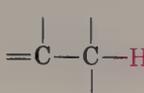
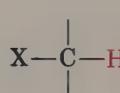
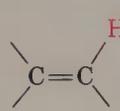
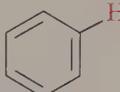
Figure 13.16. Chemical shifts for different kinds of protons.

The majority of  $^1\text{H}$  NMR absorptions occur from 0 to 8  $\delta$ , a range that can be conveniently divided into five regions, as shown in Table 13.5 (page 467). By memorizing the positions of these five regions, it's often possible to tell at a glance what general kinds of protons a molecule contains.

**TABLE 13.4** Correlation of  $^1\text{H}$  Chemical Shift with Environment

<i>Type of proton</i>	<i>Formula</i>	<i>Chemical shift (<math>\delta</math>)</i>
Reference peak	$(\text{CH}_3)_4\text{Si}$	0
Saturated primary	$-\text{CH}_3$	0.7–1.3
Saturated secondary	$-\text{CH}_2-$	1.2–1.4
Saturated tertiary	$\begin{array}{c} \diagup \\ \text{---C---H} \\ \diagdown \end{array}$	1.4–1.7
Allylic primary	$\begin{array}{c} \diagup \\ \text{C}=\text{C}-\text{CH}_3 \\ \diagdown \quad   \end{array}$	1.6–1.9
Methyl ketones	$\begin{array}{c} \text{O} \\    \\ -\text{C}-\text{CH}_3 \end{array}$	2.1–2.4
Aromatic methyl	$\text{Ar}-\text{CH}_3$	2.5–2.7
Alkyl chloride	$\begin{array}{c} \diagup \\ \text{Cl}-\text{C}-\text{H} \\ \diagdown \end{array}$	3.0–4.0
Alkyl bromide	$\begin{array}{c} \diagup \\ \text{Br}-\text{C}-\text{H} \\ \diagdown \end{array}$	2.5–4.0
Alkyl iodide	$\begin{array}{c} \diagup \\ \text{I}-\text{C}-\text{H} \\ \diagdown \end{array}$	2.0–4.0
Alcohol, ether	$\begin{array}{c} \diagup \\ -\text{O}-\text{C}-\text{H} \\ \diagdown \end{array}$	3.3–4.0
Alkynyl	$-\text{C}\equiv\text{C}-\text{H}$	2.5–2.7
Vinylic	$\begin{array}{c} \diagup \\ \text{C}=\text{C}-\text{H} \\ \diagdown \quad   \end{array}$	5.0–6.5
Aromatic	$\text{Ar}-\text{H}$	6.5–8.0
Aldehyde	$\begin{array}{c} \text{O} \\    \\ -\text{C}-\text{H} \end{array}$	9.7–10.0
Carboxylic acid	$\begin{array}{c} \text{O} \\    \\ -\text{C}-\text{O}-\text{H} \end{array}$	11.0–12.0
Alcohol	$\begin{array}{c} \diagup \\ -\text{C}-\text{O}-\text{H} \\ \diagdown \end{array}$	Extremely variable (2.5–5.0)

TABLE 13.5 Regions of the  $^1\text{H}$  NMR Spectrum

Region ( $\delta$ )	Proton type	Comments
0–1.5		Protons on carbon next to saturated centers absorb in this region. Thus, the alkane portions of most organic molecules show complex absorption here.
1.5–2.5		Protons on carbon next to unsaturated centers (allylic, benzylic, next to carbonyl) show characteristic absorptions in this region, just downfield from other alkane resonance.
2.5–4.5		Protons on carbon next to electronegative atoms (halogen, O, N) are deshielded because of the electron-withdrawing ability of these atoms. Thus, the protons absorb in this midfield region.
4.5–6.5		Protons on double-bond carbons (vinylic protons) are strongly deshielded by the neighboring pi bond and therefore absorb in this characteristic downfield region.
6.5–8.0		Protons on aromatic rings (aryl protons) are strongly deshielded by the pi orbitals of the ring and absorb in this characteristic low-field range.

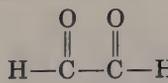
## PROBLEM.....

13.15 Each of the following compounds exhibits a single  $^1\text{H}$  NMR peak. Approximately where would you expect each compound to absorb?

(a) Cyclohexane

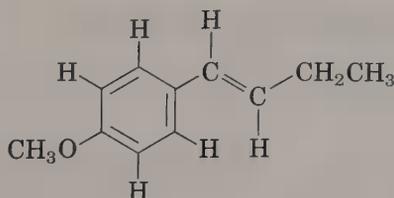
(b)  $\text{CH}_3\text{COCH}_3$ 

(c) Benzene

(d) Glyoxal (e)  $\text{CH}_2\text{Cl}_2$ (f)  $(\text{CH}_3)_3\text{N}$ 

## PROBLEM.....

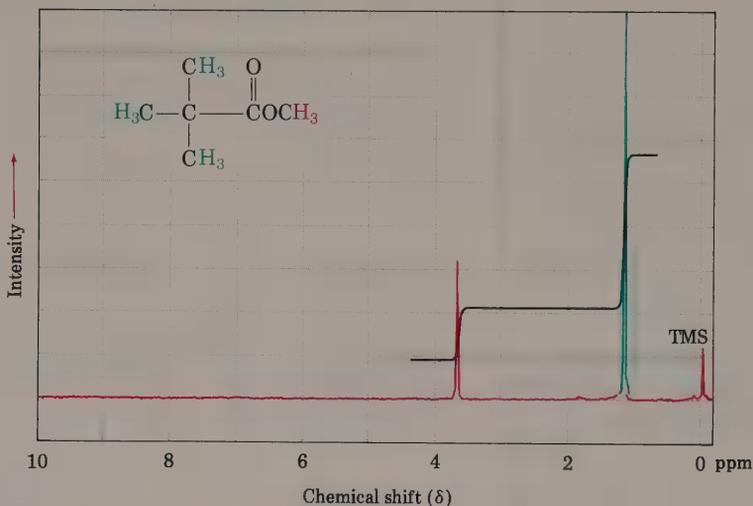
13.16 Identify the different kinds of protons in the following molecule, and tell where you would expect each to absorb.



## 13.11 Integration of $^1\text{H}$ NMR Absorptions: Proton Counting

Electronic integration of  $^1\text{H}$  NMR peak areas tells the relative numbers of protons responsible for those peaks. The idea of peak integration was introduced during the discussion of  $^{13}\text{C}$  NMR, but we remarked at that time that  $^{13}\text{C}$  integration is not normally needed because each peak is usually due to a single carbon. Integration of  $^1\text{H}$  NMR spectra, by contrast, is extremely useful because these spectra often show complicated patterns that are difficult to sort out.

An integrated  $^1\text{H}$  spectrum is displayed in a stair-step manner, with the height of each step proportional to the number of protons represented by that peak. For example, the integrated spectrum of methyl 2,2-dimethylpropanoate given in Figure 13.17 shows that the areas of the two absorptions have a ratio of 1:3. This is just what we would expect, since the three  $-\text{OCH}_3$  protons are equivalent and the nine  $(\text{CH}_3)_3\text{C}-$  protons are equivalent.



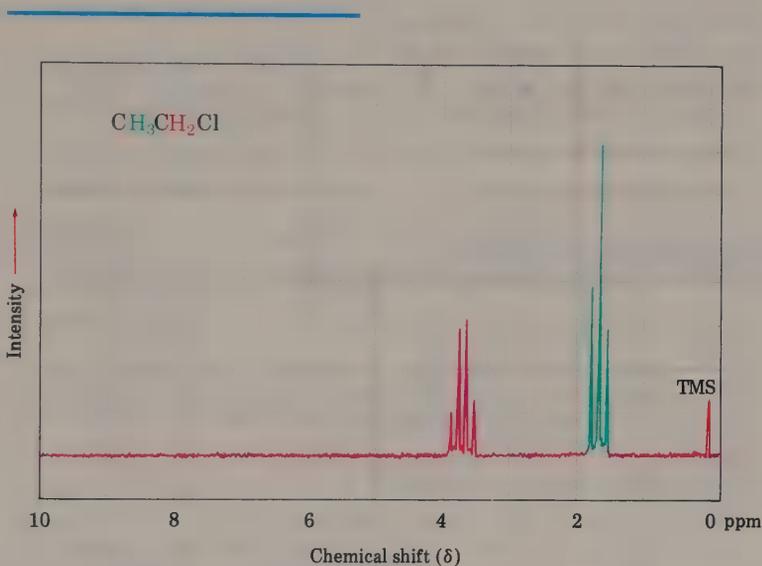
**Figure 13.17** The integrated  $^1\text{H}$  NMR spectrum of methyl 2,2-dimethylpropanoate. The two peaks have a ratio of 3:9, or 1:3.

## 13.12 Spin-Spin Splitting in $^1\text{H}$ NMR Spectra

Spin-spin splitting of single absorption peaks into multiplets is due to the coupling of neighboring nuclear spins. For example, we've seen that a  $^{13}\text{C}$  nucleus can couple with one or more nearby protons, leading to signal split-

ting when the  $^{13}\text{C}$  off-resonance NMR spectrum is recorded. Spin–spin splitting is also observed in  $^1\text{H}$  NMR.

Just as the spin of a  $^{13}\text{C}$  nucleus can couple with the spins of neighboring protons, so the spin of one proton can couple with the spins of neighboring protons. (Because of the low natural abundance of  $^{13}\text{C}$ , the coupling of a proton's spin to a  $^{13}\text{C}$  nucleus is of too low an intensity to be observed.) The resultant splitting patterns can be complex, but they can also provide much information. For example, the  $^1\text{H}$  NMR spectrum of chloroethane in Figure 13.18 shows that the  $-\text{CH}_2\text{Cl}$  protons appear as four peaks (a *quartet*) at  $3.6\ \delta$ , and that the  $-\text{CH}_3$  protons appear as a *triplet* at  $1.5\ \delta$ .

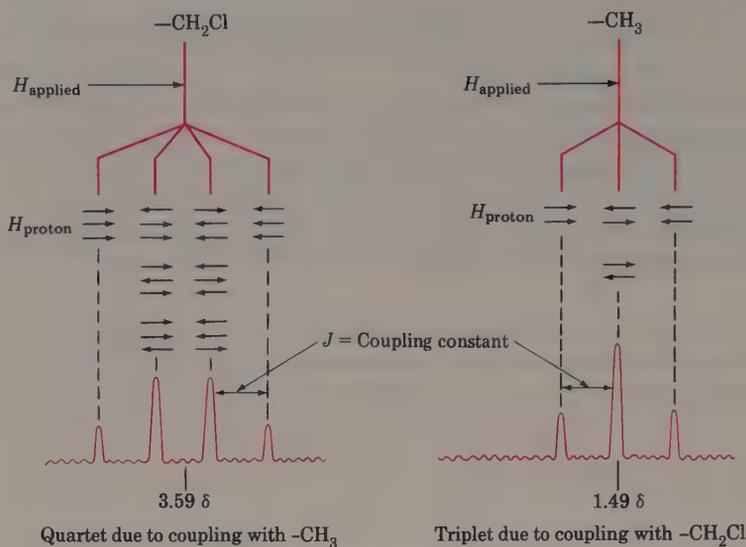


**Figure 13.18** The  $^1\text{H}$  NMR spectrum of chloroethane,  $\text{CH}_3\text{CH}_2\text{Cl}$ .

The two  $-\text{CH}_2\text{Cl}$  protons at  $3.6\ \delta$  are chemically equivalent and don't split each other's signals. Their signals *are*, however, split by the three protons on the neighboring carbon. Since each of the three neighboring protons can have its own spin aligned either with or against the applied field, there are four possible coupling combinations (remember the  $n + 1$  rule).

Similarly, the  $-\text{CH}_3$  protons of chloroethane are equivalent and don't split each other's signals. The  $-\text{CH}_3$  signals *are* split by the two neighboring  $-\text{CH}_2\text{Cl}$  protons, however, leading to the triplet observed in the  $^1\text{H}$  NMR spectrum. Figure 13.19 shows how the quartet and triplet patterns arise.

The distance between individual peaks in a multiplet is called the **coupling constant**, denoted  $J$ . Coupling constants are measured in hertz and fall in the range 0–18 Hz. The exact value of the coupling constant between two groups of protons depends on several factors, such as geometric constraints on the molecule, but a typical value for an open-chain alkyl system



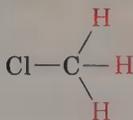
**Figure 13.19** Spin–spin splitting in chloroethane.

is 6–8 Hz. Note that a coupling constant is shared by both groups of coupled nuclei and is independent of spectrometer field strength. In chloroethane, for example, the  $\text{ClCH}_2$ -proton spins are coupled with the  $-\text{CH}_3$  proton spins and appear as a quartet with  $J = 7$  Hz. The  $-\text{CH}_3$  protons appear as a triplet with exactly the same coupling constant,  $J = 7$  Hz.

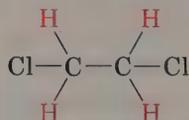
Since coupling is a reciprocal interaction between the spins of two adjacent groups of protons, we can sometimes use this fact to tell which multiplets in a complex spectrum are related to each other. It often happens that an NMR spectrum contains many multiplets, and it's sometimes difficult to tell what is coupled with what. If two multiplets have exactly the same coupling constant, however, they are probably related, and the protons causing those multiplets are therefore adjacent in the molecule.

The chloroethane spectrum in Figure 13.18 illustrates three important rules about spin–spin splitting in  $^1\text{H}$  NMR:

1. *Chemically equivalent protons do not exhibit spin–spin splitting.* The equivalent protons may be on the same carbon or on different carbons, but their signals don't split.

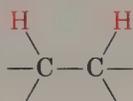


Three C–H protons are chemically equivalent; no splitting occurs.

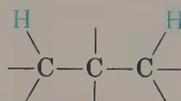


Four C–H protons are chemically equivalent; no splitting occurs.

2. The signal of a proton that has  $n$  equivalent neighboring protons is split into a multiplet of  $n + 1$  peaks with coupling constant  $J$ . Protons that are farther than two carbon atoms apart don't usually couple, although they sometimes show small coupling when they are separated by a pi bond.



Splitting observed



Splitting not usually observed

The most commonly observed coupling patterns are listed in Table 13.6, along with the relative intensities of the multiplet signals.

3. Two groups of protons coupled to each other must have the same coupling constant,  $J$ .

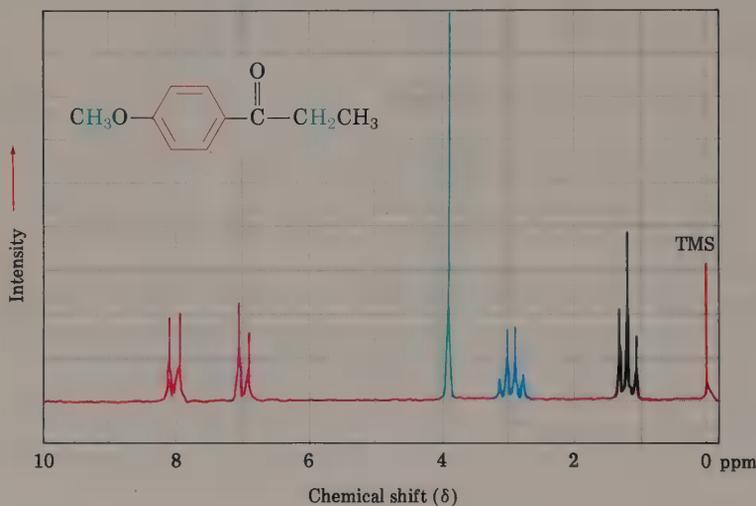
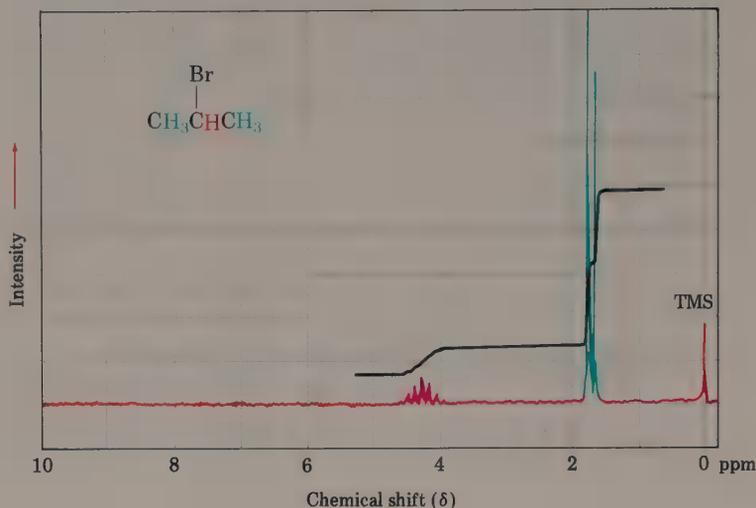
**TABLE 13.6** Some Common Spin Multiplicities

Number of equivalent adjacent protons	Type of multiplet observed	Ratio of intensities
0	Singlet	1
1	Doublet	1:1
2	Triplet	1:2:1
3	Quartet	1:3:3:1
4	Quintet	1:4:6:4:1
5	Sextet	1:5:10:10:5:1
6	Septet	1:6:15:20:15:6:1

The  $^1\text{H}$  NMR spectra of 2-bromopropane and *para*-methoxypropiophenone in Figure 13.20 further illustrate the preceding three rules.

The 2-bromopropane spectrum shows two groups of signals split into a doublet at 1.71  $\delta$  and a septet at 4.32  $\delta$ . The downfield septet is due to splitting of the  $-\text{CHBr}-$  proton signal by six equivalent neighboring protons on the two methyl groups. ( $n = 6$  leads to  $6 + 1 = 7$  peaks). The upfield doublet is due to signal splitting of the six equivalent methyl protons by the single  $-\text{CHBr}-$  proton ( $n = 1$  leads to 2 peaks). Both multiplets have the same coupling constant,  $J = 7$  Hz, and integration confirms the expected 6:1 ratio.

The *para*-methoxypropiophenone spectrum is more complex but can nevertheless be interpreted in a straightforward way. The downfield absorptions at 6.98 and 8.0  $\delta$  are due to the four aromatic ring protons. There are

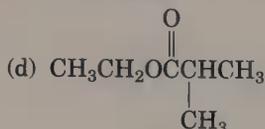
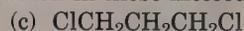
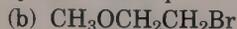


**Figure 13.20** The  $^1\text{H}$  NMR spectra of (a) 2-bromopropane and (b) *para*-methoxypropiophenone.

two kinds of protons, each of which gives a signal that is split into a doublet by its neighbor. Thus we see two doublets. The  $-\text{OCH}_3$  signal is unsplit and appears as a sharp singlet at  $3.90\ \delta$ . The  $-\text{CH}_2-$  protons next to the carbonyl group appear at  $2.95\ \delta$  in the region expected for protons on carbon next to an unsaturated center, and their signal is split into a quartet due to coupling with the neighboring methyl group. The methyl protons appear as a triplet at  $1.2\ \delta$  in the usual upfield region.

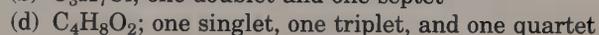
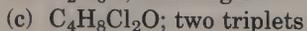
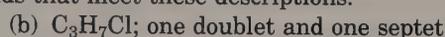
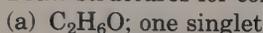
## PROBLEM

13.17 Predict the splitting patterns you would expect for each proton in these molecules.



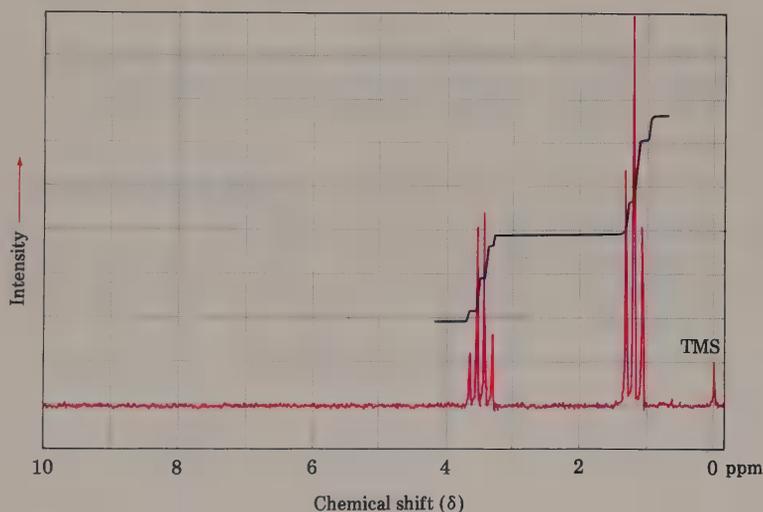
## PROBLEM

13.18 Draw structures for compounds that meet these descriptions:



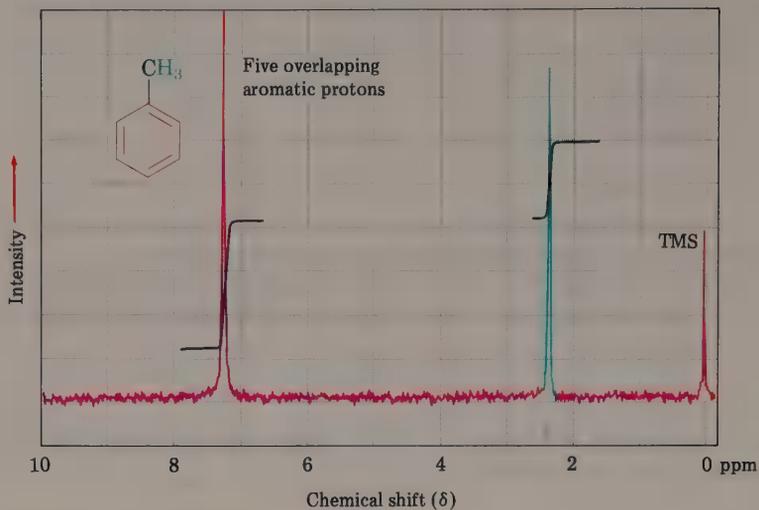
## PROBLEM

13.19 The integrated  $^1\text{H}$  NMR spectrum of a compound of formula  $\text{C}_4\text{H}_{10}\text{O}$  is shown. Propose a structure consistent with the data.



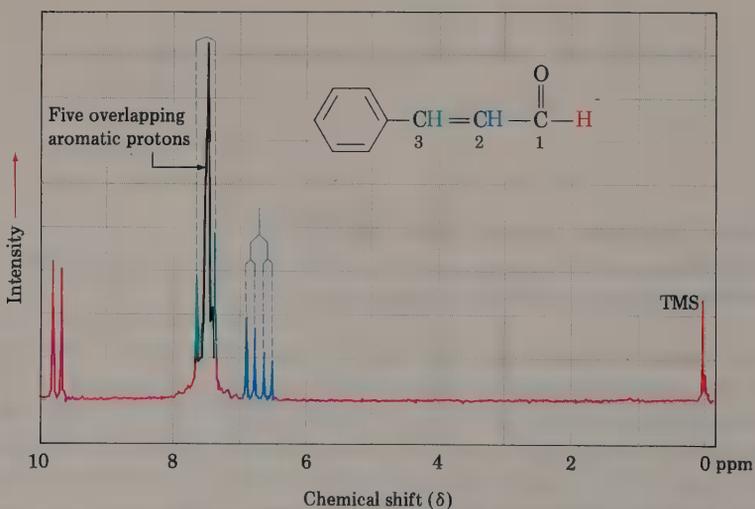
## 13.13 More Complex Spin–Spin Splitting Patterns

In all the NMR spectra we've seen so far, the chemical shifts of different protons have been quite distinct, and the spin–spin splitting patterns have been relatively simple. It often happens, however, that different kinds of protons have *overlapping* signals. The spectrum of toluene (methylbenzene) in Figure 13.21, for example, shows that all five aromatic ring protons give a single overlapping absorption, even though they aren't all equivalent. This kind of accidental overlap of signals is something we must be aware of to avoid drawing false conclusions from spectra.



**Figure 13.21** The  $^1\text{H}$  NMR spectrum of toluene, showing the accidental overlap of the nonequivalent aromatic ring protons.

Yet another complication in  $^1\text{H}$  NMR spectroscopy arises when a signal is split by two or more *nonequivalent* kinds of protons, as in the case of *trans*-cinnamaldehyde isolated from oil of cinnamon (Figure 13.22). Although the  $n + 1$  rule can be used to predict splitting due to *equivalent* protons, splittings due to nonequivalent protons are more complex.

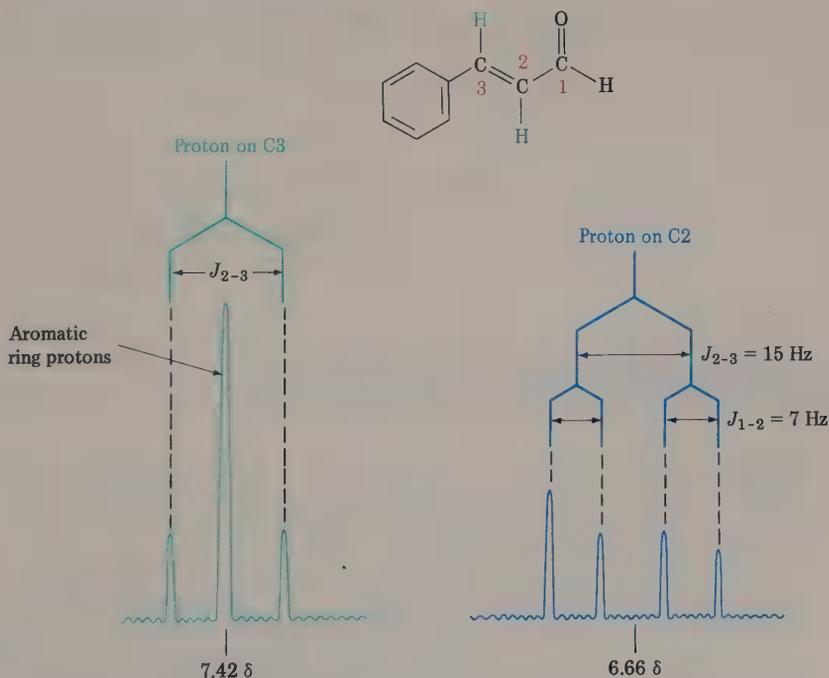


**Figure 13.22** The  $^1\text{H}$  NMR spectrum of *trans*-cinnamaldehyde.

The  $^1\text{H}$  NMR spectrum of *trans*-cinnamaldehyde can be understood if the different parts are isolated and each kind of proton is looked at individually:

1. The five aromatic proton signals overlap into a single broad resonance line at  $7.45\ \delta$ .
2. The aldehyde proton signal at C1 appears in the normal downfield position at  $9.67\ \delta$  and is split into a doublet with  $J = 7\ \text{Hz}$  by the adjacent proton at C2.
3. The vinylic proton at C3 is next to the aromatic ring and is therefore shifted downfield from the normal vinylic region. This C3 proton signal appears at  $7.42\ \delta$  and nearly overlaps the aromatic proton signals. Since it has one neighbor proton at C2, its signal is split into a doublet with  $J = 15\ \text{Hz}$ .
4. The C2 vinylic proton signal appears at  $6.66\ \delta$  and shows an interesting absorption pattern. It is coupled to the two nonequivalent protons at C1 and C3 with two different coupling constants:  $J_{1-2} = 7\ \text{Hz}$  and  $J_{2-3} = 15\ \text{Hz}$ .

The best way to see the effect of multiple coupling is to draw a tree diagram like that shown in Figure 13.23. **Tree diagrams** show the individual effects of each coupling constant on the overall pattern.



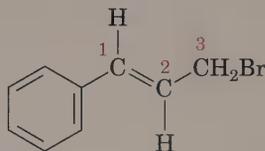
**Figure 13.23** A tree diagram for the C2 proton of *trans*-cinnamaldehyde.

The signal for the C2 proton in *trans*-cinnamaldehyde is split into a doublet by 15 Hz coupling with the C3 proton. The 7 Hz coupling with the aldehyde proton further splits each leg of the doublet into new doublets, and we therefore observe a four-line spectrum for the C2 proton. Multiple coupling can look complex but is usually amenable to simplification using tree diagrams, which consider each coupling separately.

One further point you might have noticed about the cinnamaldehyde spectrum is that the four peaks of the C2 proton resonance are not all the same size; the left-hand peaks are slightly larger. This size difference is due to the fact that the coupled nuclei have very similar chemical shifts. Whenever the chemical shifts of coupled nuclei are similar, there is a skewing of the multiplet, with the legs nearer the signal of the coupled partner becoming larger and the legs farther from the signal of the coupled partner becoming smaller. Thus, the legs of the C2 multiplet closer to the C3 absorption are larger than the legs farther away. This skewing effect on multiplets can sometimes be useful since it tells where to look in the spectrum to find the coupled partner: toward the direction of skewing.

PROBLEM .....

- 13.20 3-Bromo-1-phenyl-1-propene shows a complex NMR spectrum in which the vinylic proton at C2 is coupled with both the C1 vinylic proton ( $J = 16$  Hz) and the C3 methylene protons ( $J = 8$  Hz). Draw a tree diagram for the C2 proton signal, and account for the fact that a five-line multiplet is observed.



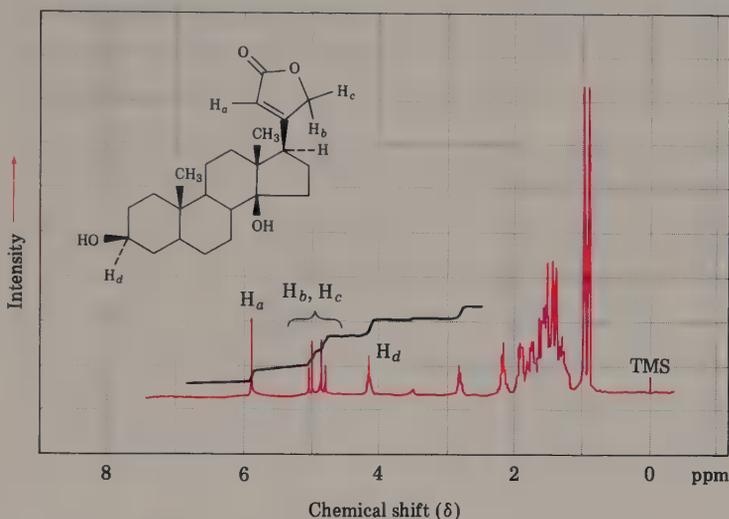
3-Bromo-1-phenyl-1-propene

## 13.14 NMR Spectra of Larger Molecules

Digitoxigenin, the steroid responsible for the heart stimulant properties of digitalis preparations, has the formula  $C_{22}H_{32}O_4$ . Can a molecule this complex possibly give an interpretable NMR spectrum?

The answer is that we often can't *fully* interpret the spectrum of such a molecule but that NMR is useful nonetheless. The  $^1\text{H}$  NMR spectrum of digitoxigenin taken on a 300 MHz instrument is shown in Figure 13.24. Even though we can't easily interpret the saturated alkane region between 1.5 and 2.0  $\delta$ , we can still recognize many structural features that provide a great deal of information. For example, we can readily spot two methyl-group singlets on saturated carbon at 0.84 and 0.92  $\delta$ . We can also see one vinylic proton singlet at 5.85  $\delta$ , a pattern at 4.85  $\delta$  due to two protons on

carbon next to oxygen, and a further broad absorption at 4.1  $\delta$ , due to one proton on carbon next to oxygen. Even in complex cases, NMR is an invaluable structural tool.

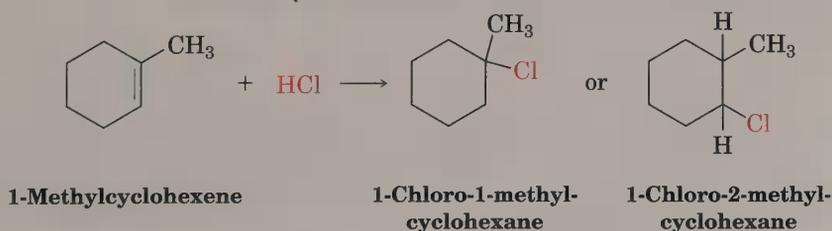


**Figure 13.24** The  $^1\text{H}$  NMR spectrum of digitoxigenin taken at 300 MHz.

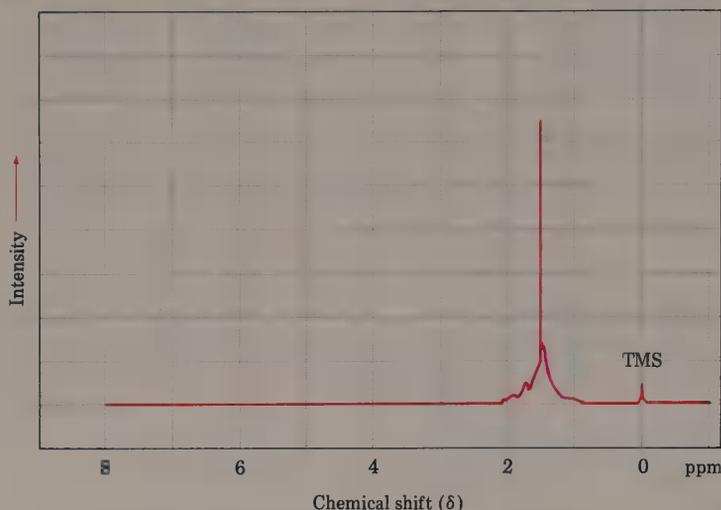
## 13.15 Uses of $^1\text{H}$ NMR Spectra

NMR can be used to help identify the product of nearly every reaction run in the laboratory. For example, we said in Chapter 7 that addition of HCl to alkenes occurs with Markovnikov regiochemistry; that is, the more highly substituted alkyl chloride is formed. With the help of NMR, we can now prove this statement.

Does addition of HCl to 1-methylcyclohexene yield 1-chloro-1-methylcyclohexane or 1-chloro-2-methylcyclohexane?



The  $^1\text{H}$  NMR spectrum of the reaction product is shown in Figure 13.25. The spectrum shows a large singlet absorption in the alkane methyl region at  $1.5\ \delta$ , indicating that the product has a methyl group bonded to a quaternary carbon,  $\text{R}_3\text{C}-\text{CH}_3$ , rather than to a tertiary carbon,  $\text{R}_2\text{CH}-\text{CH}_3$ . Furthermore, the spectrum shows *no* absorptions in the range  $4\text{--}5\ \delta$ , where we would expect the signal of an  $\text{R}_2\text{CHCl}$  proton to occur. Thus, it's clear that 1-chloro-1-methylcyclohexane is the reaction product.



**Figure 13.25** The  $^1\text{H}$  NMR spectrum of the reaction product from  $\text{HCl}$  and 1-methylcyclohexene.

**PROBLEM** .....

- 13.21 How could you use  $^1\text{H}$  NMR to help you determine the regiochemistry of alkene hydroboration? Does hydroboration–oxidation of propene yield 1-propanol or 2-propanol?
- .....

## 13.16 Summary and Key Words

When  $^1\text{H}$  and  $^{13}\text{C}$  nuclei are placed in strong magnetic fields, their spins orient either with or against the applied field. On irradiation with radio frequency (rf) waves, energy is absorbed and the nuclei “spin-flip” from the lower energy state to the higher energy state. This absorption of rf energy is detected, amplified, and displayed as a **nuclear magnetic resonance (NMR) spectrum**.

The NMR spectrum is obtained by irradiating a sample with a constant-frequency rf energy and slowly changing the value of the applied magnetic field. Different kinds of  $^1\text{H}$  and  $^{13}\text{C}$  nuclei come into resonance at slightly different applied fields, and we therefore see a different absorption line for each different kind of  $^1\text{H}$  and  $^{13}\text{C}$ . The NMR chart is calibrated in **delta** ( $\delta$ ) **units**, where  $1 \delta = 1$  part per million (ppm) of spectrometer frequency. Tetramethylsilane (TMS) is used as a reference point to which other peaks are compared because it shows both  $^1\text{H}$  and  $^{13}\text{C}$  absorptions at unusually high values of the applied magnetic field. The TMS absorption occurs at the right-hand (**upfield**) edge of the chart and is arbitrarily assigned a value of  $0 \delta$ .

Both  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra display four general features:

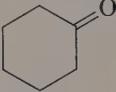
1. *Number of resonance lines* Each different kind of  $^1\text{H}$  or  $^{13}\text{C}$  nucleus in a molecule can give rise to a different resonance line.
2. *Chemical shift* The exact position of each peak is its **chemical shift**. Chemical shifts are due to the effects of electrons setting up tiny local magnetic fields that **shield** a nearby nucleus from the applied field and therefore cause different nuclei to come to resonance at different places. By correlating chemical shifts with environment, we can learn about the chemical nature of each nucleus.
3. *Integration* The area under each NMR absorption peak can be electronically **integrated** to determine the relative number of nuclei ( $^1\text{H}$  or  $^{13}\text{C}$ ) responsible for each peak.
4. *Spin-spin splitting* Neighboring nuclear spins can **couple**, splitting NMR peaks into **multiplets**. The NMR signal of a  $^{13}\text{C}$  nucleus bonded to  $n$  protons splits into  $n + 1$  peaks (the  **$n + 1$  rule**). Similarly, the NMR signal of a  $^1\text{H}$  nucleus neighbored by  $n$  equivalent adjacent protons splits into  $n + 1$  peaks.

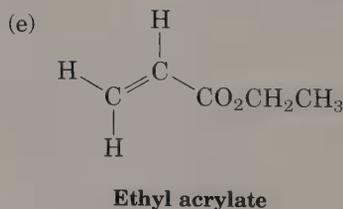
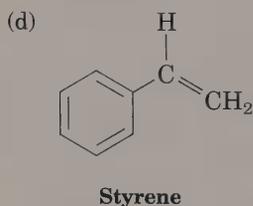
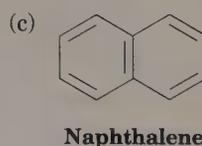
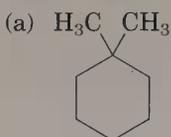
Most  $^{13}\text{C}$  spectra are run in an operating mode that provides maximum sensitivity and gives a spectrum in which each nonequivalent carbon shows a single unsplit resonance line. Operating in the integrating mode causes a loss of sensitivity but provides a spectrum that can be electronically integrated to measure the number of carbon nuclei responsible for each peak. Operating in the spin-coupled mode also causes a loss of sensitivity but provides a spectrum in which spin-spin splitting is observed. Each carbon resonance is split into a multiplet depending on the number of protons to which it is bonded: Quaternary carbon resonances remain as singlets, tertiary carbons ( $\text{R}_3\text{CH}$ ) appear as doublets, secondary carbons ( $\text{R}_2\text{CH}_2$ ) appear as triplets, and primary carbons ( $\text{RCH}_3$ ) appear as quartets.

Proton NMR spectra are even more useful than  $^{13}\text{C}$  spectra. The sensitivity of  $^1\text{H}$  instruments is high, and normal spectrometer operating conditions provide spectra that show spin-spin splitting and can be integrated. Proton resonances usually fall into the range  $0$ – $10 \delta$  downfield from the TMS reference point.

Specific  $^1\text{H}$  resonance peaks are often split into multiplets due to spin-spin splitting with the spins of protons on adjacent carbons. Equivalent protons don't split each other, but a proton with  $n$  equivalent neighboring protons gives a signal that is split into  $n + 1$  peaks with **coupling constant  $J$** .

## ADDITIONAL PROBLEMS .....

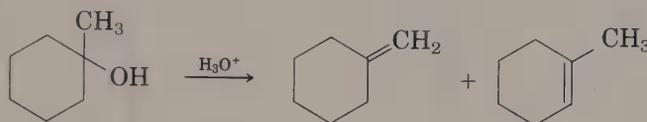
- 13.22** The following  $^1\text{H}$  NMR absorptions were determined on a spectrometer operating at 60 MHz and are given in hertz downfield from the TMS standard. Convert the absorptions to  $\delta$  units.  
 (a) 131 Hz                      (b) 287 Hz                      (c) 451 Hz                      (d) 543 Hz
- 13.23** The following  $^1\text{H}$  NMR absorptions given in  $\delta$  units were obtained on a spectrometer operating at 80 MHz. Convert the chemical shifts from  $\delta$  units into hertz downfield from TMS.  
 (a) 2.1                      (b) 3.45                      (c) 6.30                      (d) 7.70
- 13.24** When measured on a spectrometer operating at 60 MHz, chloroform ( $\text{CHCl}_3$ ) shows a single sharp absorption at 7.3  $\delta$ .  
 (a) How many parts per million downfield from TMS does chloroform absorb?  
 (b) How many hertz downfield from TMS would chloroform absorb if the measurement were carried out on a spectrometer operating at 360 MHz?  
 (c) What would be the position of the chloroform absorption in  $\delta$  units when measured on a 360 MHz spectrometer?
- 13.25** How many absorptions would you expect to observe in the normal  $^{13}\text{C}$  NMR spectra of these compounds?  
 (a) 1,1-Dimethylcyclohexane                      (b)  $\text{CH}_3\text{CH}_2\text{OCH}_3$   
 (c) *tert*-Butylcyclohexane                      (d) 3-Methyl-1-pentene  
 (e) *cis*-1,2-Dimethylcyclohexane                      (f) 
- 13.26** Indicate the spin multiplicities you would expect to see for each carbon atom in the spin-coupled  $^{13}\text{C}$  NMR spectra of the molecules shown in Problem 13.25.
- 13.27** Why do you suppose accidental overlap of signals is much more common in  $^1\text{H}$  NMR than in  $^{13}\text{C}$  NMR?
- 13.28** Tell what is meant by each of these terms:  
 (a) Chemical shift                      (b) Spin-spin splitting  
 (c) Applied magnetic field                      (d) Spectrometer operating frequency  
 (e) Coupling constant                      (f) Upfield/downfield
- 13.29** How many types of nonequivalent protons are there in each of the following molecules?



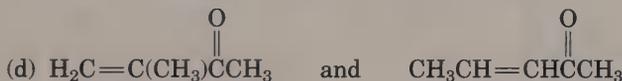
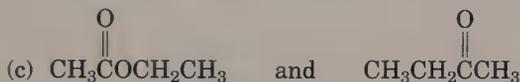
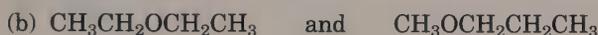
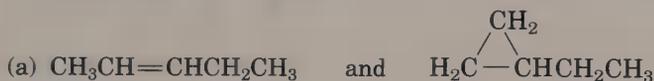
- 13.30** The following compounds all show a single line in their  $^1\text{H}$  NMR spectra. List them in expected order of increasing chemical shift.



- 13.31 Predict the splitting pattern for each kind of hydrogen and each kind of carbon in these molecules:  
 (a)  $(\text{CH}_3)_3\text{CH}$                       (b)  $\text{CH}_3\text{CH}_2\text{COOCH}_3$                       (c) *trans*-2-Butene
- 13.32 Predict the splitting pattern for each kind of hydrogen and each kind of carbon in isopropyl propanoate,  $\text{CH}_3\text{CH}_2\text{COOCH}(\text{CH}_3)_2$ .
- 13.33 The acid-catalyzed dehydration of 1-methylcyclohexanol yields a mixture of two alkenes as product. After you separated them by chromatography, how would you use  $^1\text{H}$  NMR to help you decide which was which?



- 13.34 How would you use  $^1\text{H}$  NMR to distinguish between these pairs of isomers?



- 13.35 Propose structures for compounds with the following formulas that show only one peak in their  $^1\text{H}$  NMR spectra.



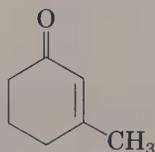
- 13.36 Assume that you have a compound with formula  $\text{C}_3\text{H}_6\text{O}$ .

- (a) How many double bonds and/or rings does your material contain?  
 (b) Propose as many structures as you can that fit the molecular formula.  
 (c) If your compound shows an infrared absorption peak at  $1715\text{ cm}^{-1}$ , what inferences can you draw?  
 (d) If your compound shows a single  $^1\text{H}$  NMR absorption peak at  $2.1\ \delta$ , what is its structure?

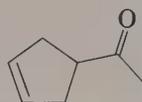
- 13.37 How would you use  $^1\text{H}$  and  $^{13}\text{C}$  NMR to help you distinguish among the following isomeric compounds of formula  $\text{C}_4\text{H}_8$ ?



- 13.38 How could you use  $^1\text{H}$  and  $^{13}\text{C}$  NMR to help you distinguish between these structures?

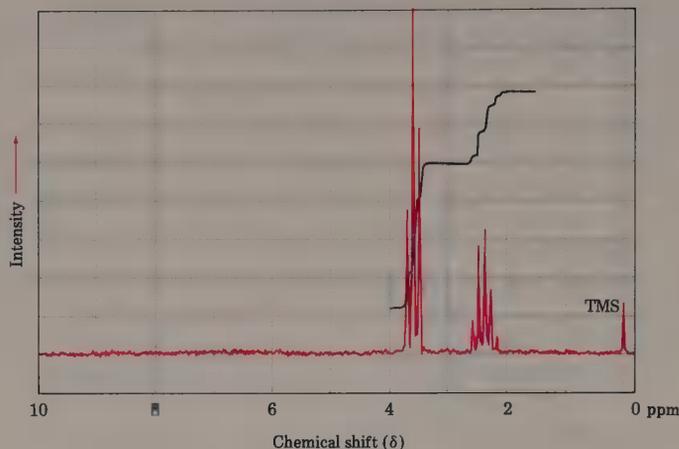


3-Methyl-2-cyclohexenone

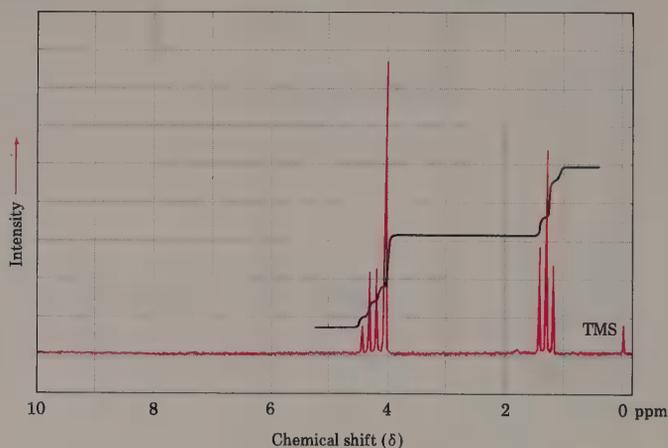


4-Cyclopentenyl methyl ketone

- 13.39 How could you use IR spectroscopy to help you distinguish between the two compounds shown in Problem 13.38?
- 13.40 The compound whose  $^1\text{H}$  NMR spectrum is shown here has the molecular formula  $\text{C}_3\text{H}_6\text{Br}_2$ . Propose a plausible structure.



- 13.41 Propose structures for compounds that fit the following  $^1\text{H}$  NMR data:
- |   |                                     |
|---|-------------------------------------|
| (a) $\text{C}_5\text{H}_{10}\text{O}$     | (b) $\text{C}_3\text{H}_5\text{Br}$ |
| 6 H doublet at 0.95 $\delta$ , $J = 7$ Hz | 3 H singlet at 2.32 $\delta$        |
| 3 H singlet at 2.10 $\delta$              | 1 H broad singlet at 5.35 $\delta$  |
| 1 H multiplet at 2.43 $\delta$            | 1 H broad singlet at 5.54 $\delta$  |
- 13.42 The compound whose  $^1\text{H}$  NMR spectrum is shown has the molecular formula  $\text{C}_4\text{H}_7\text{O}_2\text{Cl}$  and shows an infrared absorption peak at  $1740\text{ cm}^{-1}$ . Propose a plausible structure.



- 13.43 Propose structures for compounds that fit the following  $^1\text{H}$  NMR data:
- |   |                                  |
|---|----------------------------------|
| (a) $\text{C}_4\text{H}_8\text{Cl}_2$     | (b) $\text{C}_{10}\text{H}_{14}$ |
| 3 H singlet at 2.18 $\delta$              | 9 H singlet at 1.30 $\delta$     |
| 2 H doublet at 4.16 $\delta$ , $J = 7$ Hz | 5 H singlet at 7.30 $\delta$     |
| 1 H triplet at 5.71 $\delta$ , $J = 7$ Hz |                                  |

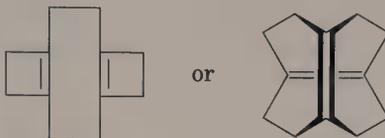
(c)  $C_4H_7BrO$ 

3 H singlet at 2.11  $\delta$   
 2 H triplet at 3.52  $\delta$ ,  $J = 6$  Hz  
 2 H triplet at 4.40  $\delta$ ,  $J = 6$  Hz

(d)  $C_9H_{11}Br$ 

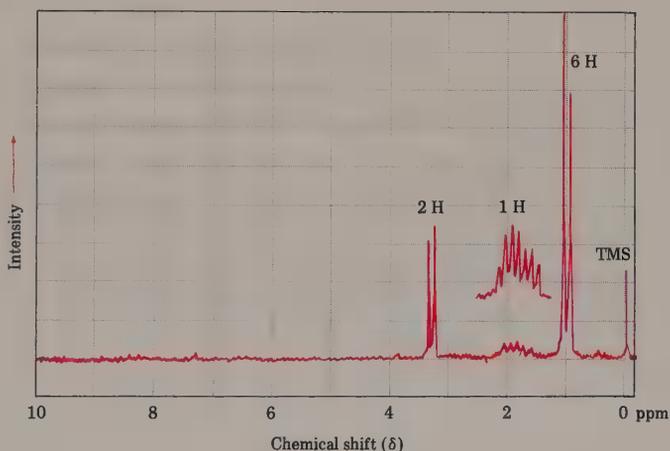
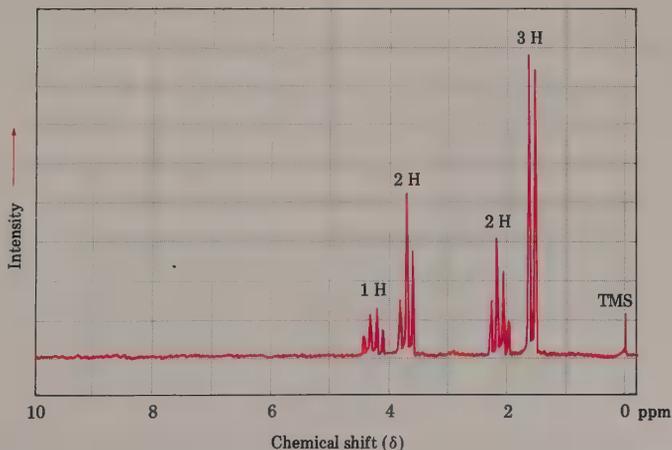
2 H quintet at 2.15  $\delta$ ,  $J = 7$  Hz  
 2 H triplet at 2.75  $\delta$ ,  $J = 7$  Hz  
 2 H triplet at 3.38  $\delta$ ,  $J = 7$  Hz  
 5 H singlet at 7.22  $\delta$

13.44 How might you use NMR (either  $^1H$  or  $^{13}C$ ) to differentiate between the following two isomeric structures?



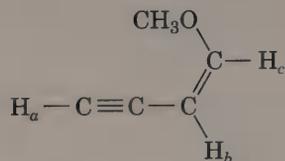
(You might want to build molecular models to help you examine the two structures more closely.)

13.45 Propose plausible structures for the two compounds whose  $^1H$  NMR spectra are shown.

(a)  $C_4H_9Br$ (b)  $C_4H_8Cl_2$ 

13.46 We saw earlier that long-range coupling between protons more than two carbon atoms apart is sometimes observed when pi bonds intervene. One example of long-

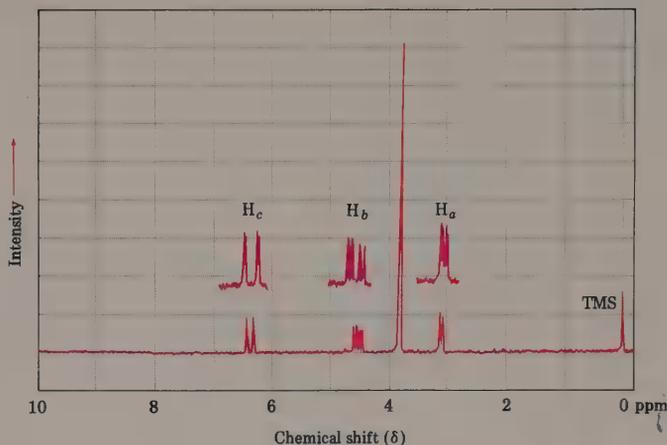
range coupling is found in 1-methoxy-1-buten-3-yne, whose  $^1\text{H}$  NMR spectrum is shown.



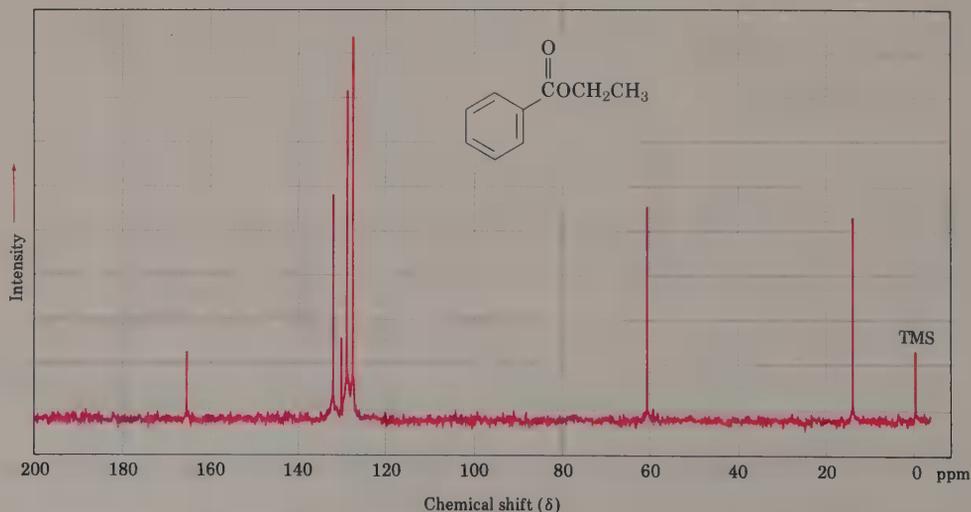
Not only does the acetylenic proton,  $\text{H}_a$ , couple with the vinylic proton  $\text{H}_b$ , it also couples with the vinylic proton  $\text{H}_c$  (four carbon atoms away). The following coupling constants are observed:

$$J_{a-b} = 3 \text{ Hz} \quad J_{a-c} = 1 \text{ Hz} \quad J_{b-c} = 7 \text{ Hz}$$

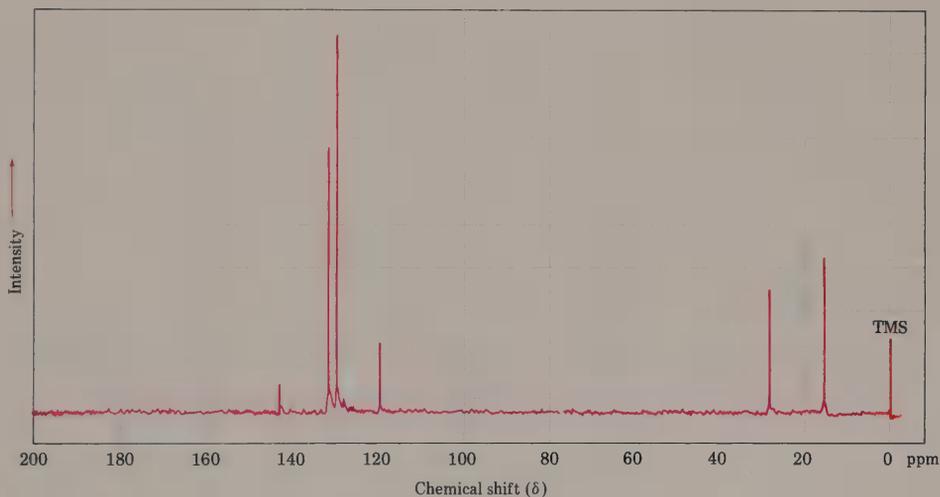
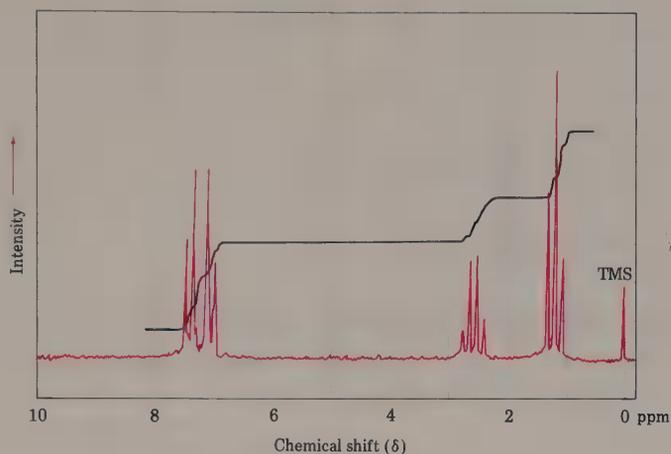
Construct tree diagrams that account for the observed splitting patterns of  $\text{H}_a$ ,  $\text{H}_b$ , and  $\text{H}_c$ .



- 13.47 Assign as many of the resonances as you can to specific carbon atoms in the  $^{13}\text{C}$  NMR spectrum of ethyl benzoate shown here.

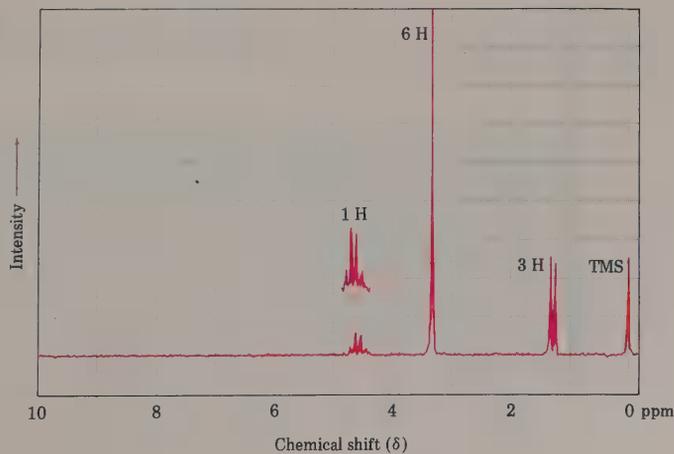


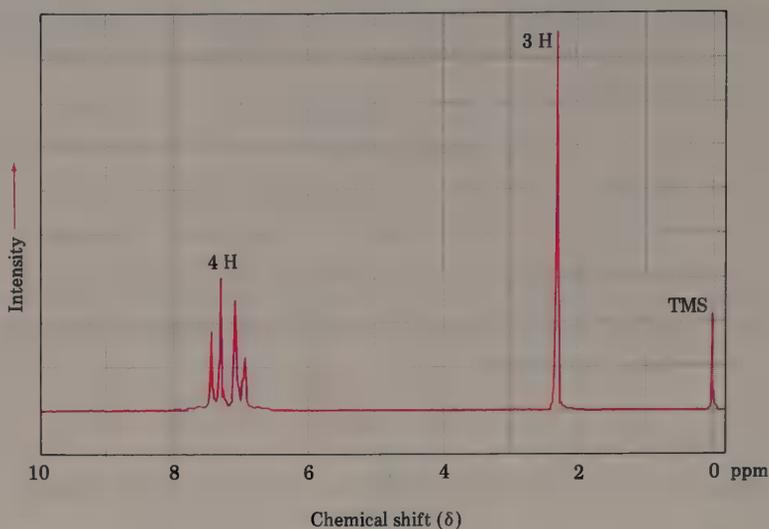
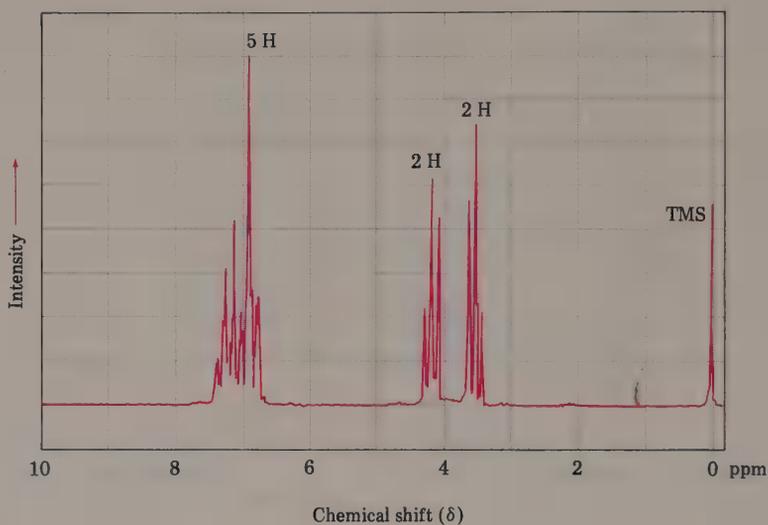
- 13.48 The  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of compound A,  $\text{C}_8\text{H}_9\text{Br}$ , are shown. Propose a possible structure for A, and assign peaks in the spectra to your structure.



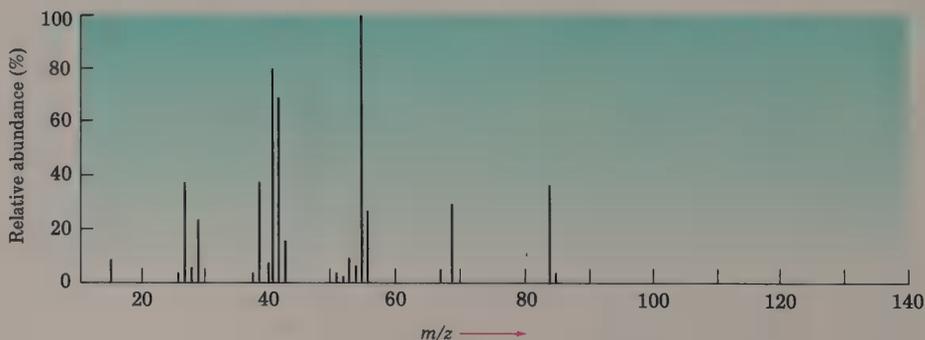
- 13.49 Propose plausible structures for the three compounds whose  $^1\text{H}$  NMR spectra are shown.

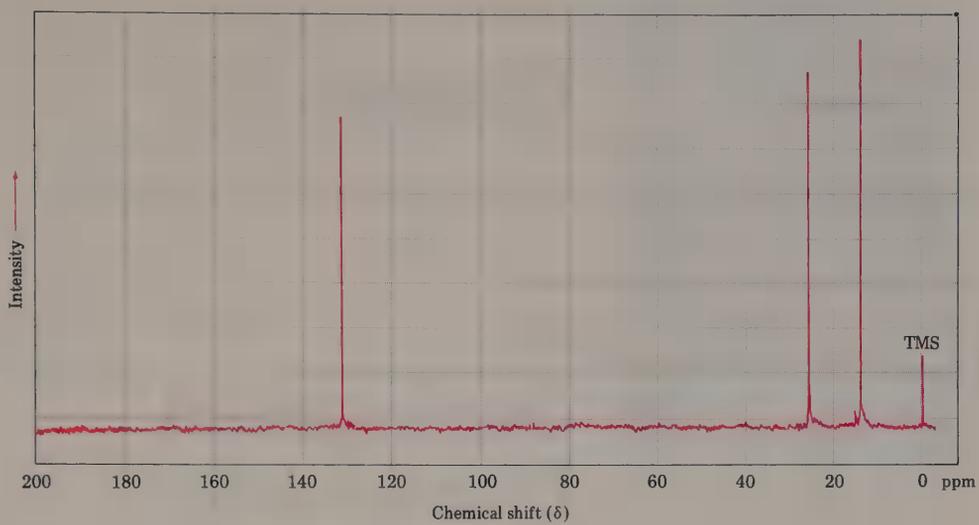
(a)  $\text{C}_4\text{H}_{10}\text{O}_2$

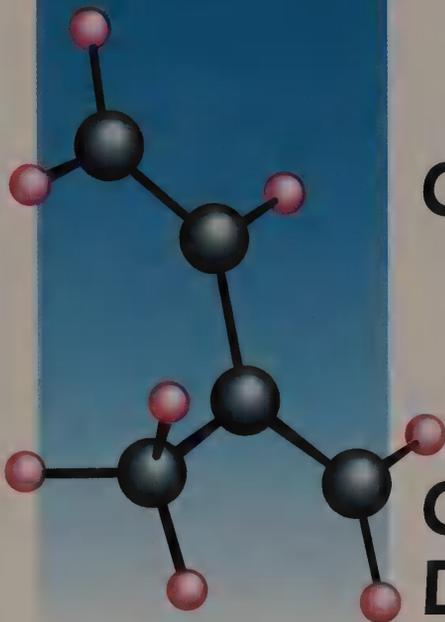


(b)  $C_7H_7Br$ (c)  $C_8H_9Br$ 

**13.50** The mass spectrum and  $^{13}C$  NMR spectrum of a hydrocarbon are shown. Propose a suitable structure for this hydrocarbon, and explain the spectral data.



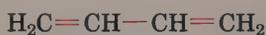




## Chapter 14

# Conjugated Dienes and Ultraviolet Spectroscopy

Double bonds that alternate with single bonds are said to be **conjugated**. Thus, 1,3-butadiene is a **conjugated diene**, whereas 1,4-pentadiene is a nonconjugated diene with *isolated* double bonds.



**1,3-Butadiene**

(conjugated; alternating  
double and single bonds)



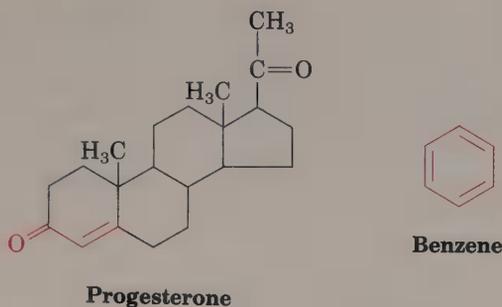
**1,4-Pentadiene**

(nonconjugated; nonalternating  
double and single bonds)

There are other types of conjugated systems besides dienes, many of which play important roles in nature and biology. For example, the pigments responsible for the brilliant reds and yellows of fruits and flowers are conjugated *polyenes* (*poly* = “many”). Lycopene, the red pigment in tomatoes, is one such molecule. Conjugated *enones* (alkene + ketone) are common structural features of important molecules such as progesterone, the so-called pregnancy hormone. Conjugated cyclic molecules such as benzene are a major field of study in themselves and will be considered in detail in the next chapter.

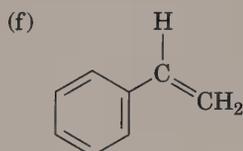
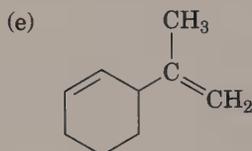
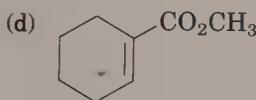
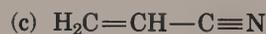
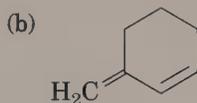
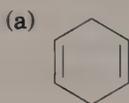


**Lycopene**



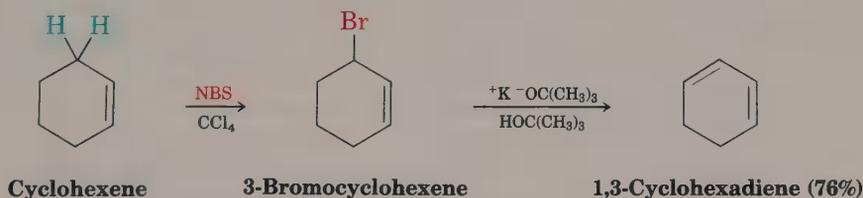
## PROBLEM

14.1 Which of the following molecules contain conjugated systems? Circle the conjugated portion.

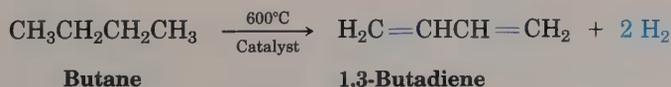


## 14.1 Preparation of Conjugated Dienes

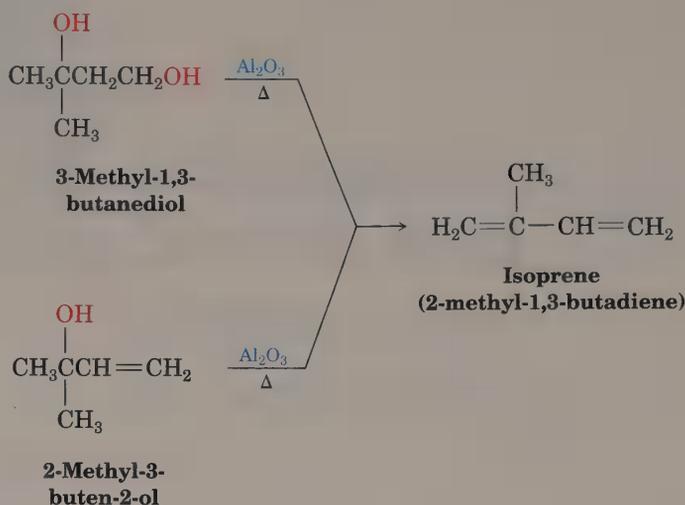
Conjugated dienes are generally prepared by the methods previously discussed for alkene synthesis. For example, the base-induced elimination of HX from an allylic halide produces a conjugated diene.



1,3-Butadiene itself is prepared industrially for use in polymer synthesis. One industrial method involves thermal cracking of butane over a special chromium oxide–aluminum oxide catalyst, but this procedure is of no use in the laboratory.



Other simple conjugated dienes that have important uses in polymer synthesis include isoprene (2-methyl-1,3-butadiene) and chloroprene (2-chloro-1,3-butadiene). Isoprene has been prepared industrially by a number of methods, including the dehydration of 2-methyl-3-buten-2-ol and the double dehydration of 3-methyl-1,3-butanediol over an alumina catalyst, but these methods are rarely used in the laboratory.



## 14.2 Stability of Conjugated Dienes

Conjugated dienes are similar to isolated alkenes in much of their chemistry. There are, however, some important differences such as *stability*. Conjugated dienes are somewhat more stable than nonconjugated dienes.

Evidence for the extra stability of conjugated dienes comes from measurements of heats of hydrogenation (Table 14.1). We saw earlier in the discussion of alkene stabilities (Section 6.7) that alkenes of similar substitution pattern have remarkably similar  $\Delta H^\circ_{\text{hydrog}}$  values. Monosubstituted alkenes such as 1-butene have values for  $\Delta H^\circ_{\text{hydrog}}$  near 30 kcal/mol (125 kJ/mol), whereas disubstituted alkenes such as 2-methyl-1-butene show  $\Delta H^\circ_{\text{hydrog}}$  values approximately 3 kcal/mol lower. We concluded from these data that highly substituted alkenes are more stable than less highly substituted ones. That is, substituted alkenes release less heat on hydrogenation because they contain less energy to start with. A similar conclusion can be drawn for conjugated dienes.

Since a monosubstituted alkene such as 1-butene has  $\Delta H^\circ_{\text{hydrog}} = 30.3$  kcal/mol (127 kJ/mol), we might predict that a compound with two monosubstituted double bonds should have a  $\Delta H^\circ_{\text{hydrog}}$  approximately twice this value, or 60.6 kcal/mol. This prediction is met by nonconjugated dienes such as 1,4-pentadiene ( $\Delta H^\circ_{\text{hydrog}} = 60.8$  kcal/mol) but is *not* met by the conjugated diene 1,3-butadiene ( $\Delta H^\circ_{\text{hydrog}} = 57.1$  kcal/mol). 1,3-Butadiene is approximately 3.7 kcal/mol (15 kJ/mol) more stable than predicted.

TABLE 14.1 Heats of Hydrogenation for Some Alkenes and Dienes

Alkene	Product	$\Delta H^\circ_{\text{hydrog}}$	
		(kcal/mol)	(kJ/mol)
$\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$ <b>1-Butene</b>	$\xrightarrow{\text{H}_2}$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	30.3	127
$\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3\text{CH}_2\text{C}=\text{CH}_2 \end{array}$ <b>2-Methyl-1-butene</b>	$\xrightarrow{\text{H}_2}$ $\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3\text{CH}_2\text{CHCH}_3 \end{array}$	26.9	113
$\text{H}_2\text{C}=\text{CHCH}=\text{CH}_2$ <b>1,3-Butadiene</b>	$\xrightarrow{\text{H}_2}$ $\text{CH}_3\text{CH}_2\text{CH}=\text{CH}_2$	26.7	112
$\text{H}_2\text{C}=\text{CHCH}=\text{CH}_2$ <b>1,3-Butadiene</b>	$\xrightarrow{2 \text{H}_2}$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$	57.1	239
$\begin{array}{c} \text{CH}_3 \\   \\ \text{H}_2\text{C}=\text{CHC}=\text{CH}_2 \end{array}$ <b>2-Methyl-1,3-butadiene</b>	$\xrightarrow{2 \text{H}_2}$ $\begin{array}{c} \text{CH}_3 \\   \\ \text{CH}_3\text{CH}_2\text{CHCH}_3 \end{array}$	53.4	223
$\text{H}_2\text{C}=\text{CHCH}_2\text{CH}=\text{CH}_2$ <b>1,4-Pentadiene</b>	$\xrightarrow{2 \text{H}_2}$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	60.8	254
$\text{H}_2\text{C}=\text{CHCH}_2\text{CH}_2\text{CH}=\text{CH}_2$ <b>1,5-Hexadiene</b>	$\xrightarrow{2 \text{H}_2}$ $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$	60.5	253

Confirmation of this unexpected stability comes from data on the partial hydrogenation of 1,3-butadiene. If 1,3-butadiene is partially hydrogenated to yield 1-butene, 26.7 kcal/mol (112 kJ/mol) energy is released, a value 3.6 kcal/mol less than we would expect for a normal isolated monosubstituted double bond. The same is true of other conjugated dienes.

	$\Delta H^\circ_{\text{hydrog}}$ (kcal/mol)	
$\text{H}_2\text{C}=\text{CHCH}_2\text{CH}=\text{CH}_2$	30.3 + 30.3 =	60.6 Expected
<b>1,4-Pentadiene</b>		60.8 Observed
		-0.2 Difference
$\text{H}_2\text{C}=\text{CHCH}=\text{CH}_2$	30.3 + 30.3 =	60.6 Expected
<b>1,3-Butadiene</b>		57.1 Observed
		3.5 Difference
$\begin{array}{c} \text{CH}_3 \\   \\ \text{H}_2\text{C}=\text{CCH}=\text{CH}_2 \end{array}$	30.3 + 26.9 =	57.2 Expected
<b>2-Methyl-1,3-butadiene</b>		53.4 Observed
		3.8 Difference

## PROBLEM .....

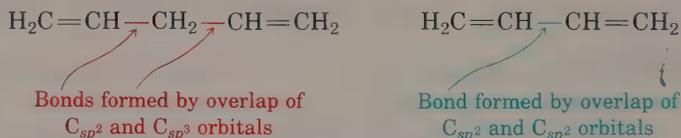
- 14.2 Use the data in Table 14.1 to calculate an expected heat of hydrogenation for allene,  $\text{H}_2\text{C}=\text{C}=\text{CH}_2$ . The measured value is 71.3 kcal/mol (298 kJ/mol). How stable is allene? Rank a conjugated diene, a nonconjugated diene, and an allene in order of stability.
- .....

### 14.3 Molecular Orbital Description of 1,3-Butadiene

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Why are conjugated dienes unusually stable? Two different explanations have been advanced. One explanation says that 1,3-butadiene really does *not* have “unusual” stability; the problem is with what we call “usual,” not with the diene. According to this view, our expectations are wrong because a nonconjugated diene such as 1,4-pentadiene is a poor choice for comparison with 1,3-butadiene.

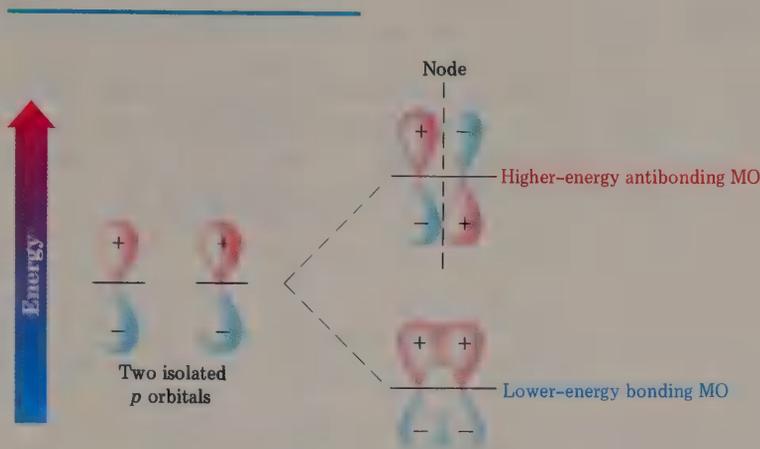
In nonconjugated dienes, the carbon–carbon single bonds result from sigma overlap of an  $sp^2$  orbital from one carbon with an  $sp^3$  orbital from the neighboring carbon. In conjugated dienes, however, the carbon–carbon single bonds result from sigma overlap of  $sp^2$  orbitals on both carbons. Since  $sp^2$  orbitals have more  $s$  character than  $sp^3$  orbitals do, they form somewhat shorter, stronger bonds. Thus, the “extra” stability of a conjugated diene has more to do with the hybridization of the orbitals forming the carbon–carbon single bond than with the double bonds themselves.



The second explanation for the stability of conjugated dienes focuses on the interaction between pi orbitals of the conjugated diene system. To see how this interaction of pi orbitals arises, let’s briefly review molecular orbital theory (Section 1.7). When a covalent bond is formed by overlap of two atomic orbitals, the orbitals that result are the property of the molecule, not of the individual atoms. The electrons in those orbitals are shared between atoms, rather than localized on one atom. For example, when two  $p$  atomic orbitals overlap in a sideways manner to form a pi bond, the bonding electrons occupy a pi *molecular* orbital, rather than  $p$  atomic orbitals.

Since two  $p$  atomic orbitals are involved in forming a pi bond, two pi molecular orbitals are formed. One is lower in energy than the starting  $p$  orbitals and is therefore a bonding molecular orbital; the other is higher in energy and is an antibonding orbital. When we assign electrons to the orbitals, both electrons go into the low-energy bonding orbital, resulting in formation of a stable bond (Figure 14.1).

The pictorial representation of molecular orbital formation in Figure 14.1 shows how bonding arises. We saw earlier that  $p$  orbitals are dumbbell



**Figure 14.1** Two isolated  $p$  orbitals combine to form two pi molecular orbitals. When these orbitals are filled, both electrons occupy the low-energy bonding orbital, leading to a net lowering of energy and formation of a stable bond.

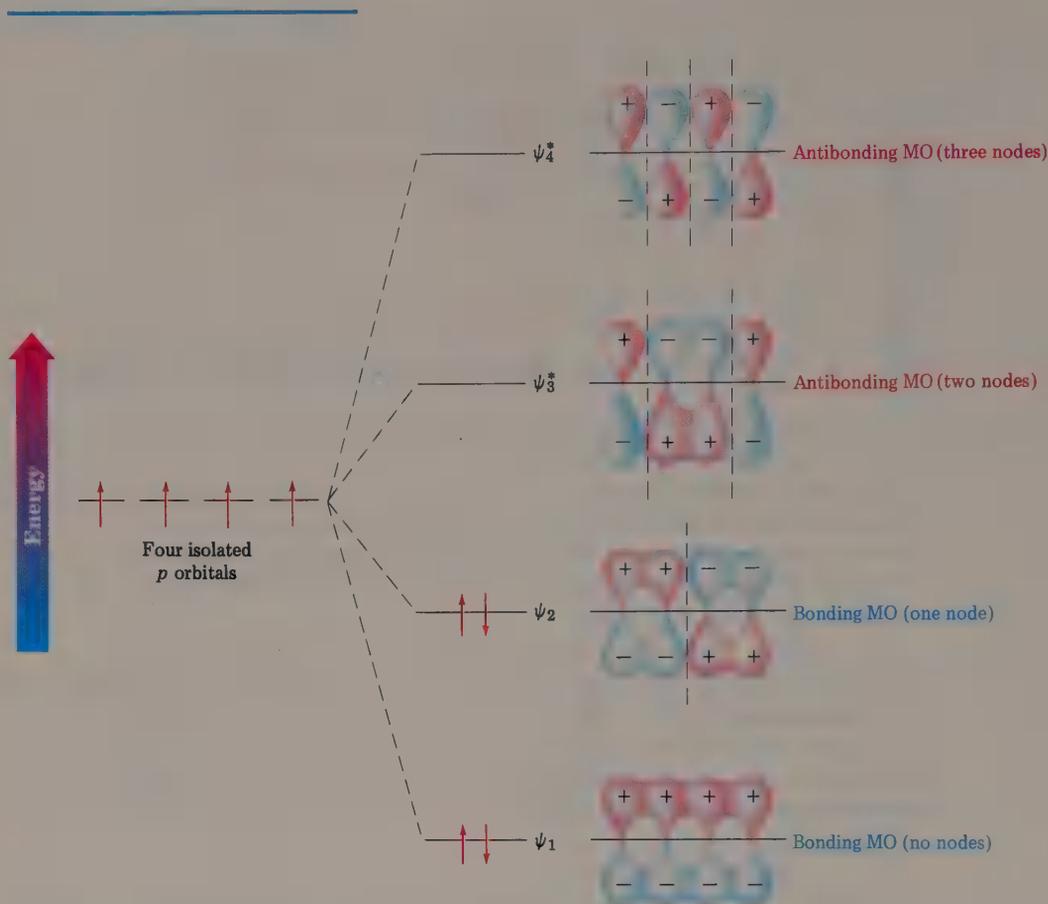
shaped and that the two lobes have different mathematical signs. When two  $p$  orbitals approach each other for overlap, they can orient in either of two ways. Overlap of lobes with identical signs is additive and corresponds to the low-energy bonding molecular orbital, whereas overlap of lobes with different signs is subtractive and corresponds to the high-energy antibonding orbital.

The change of sign between adjacent lobes in the antibonding orbital means that, if there were electrons in this orbital, they would not be shared between nuclei and there would be no bonding. Thus, there is a *node*, or region of zero electron density between the two nuclei. The bonding molecular orbital has no node between the nuclei. (There is also a node between the plus and minus lobes of each individual pi orbital, but only the nodes between nuclei affect bonding.)

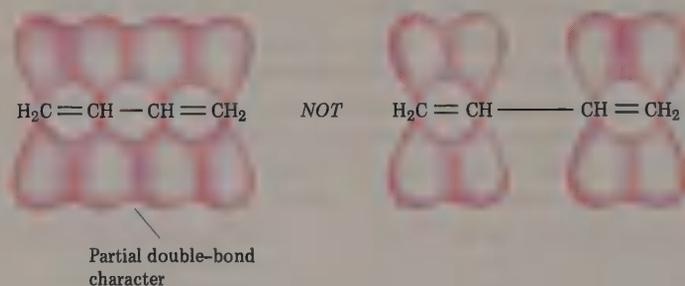
Now let's bring four  $p$  atomic orbitals together and allow two pi bonds to form, as occurs in a conjugated diene. In so doing, we generate a set of four pi molecular orbitals (Figure 14.2).

The lowest-energy molecular orbital (denoted  $\psi_1$ , Greek psi) is a fully additive combination that has no nodes between the nuclei. It is therefore a bonding orbital and holds two electrons. The molecular orbital of next lowest energy,  $\psi_2$ , has one node between nuclei, is also a bonding orbital, and holds the remaining two electrons. Above  $\psi_1$  and  $\psi_2$  in energy are the two antibonding molecular orbitals,  $\psi_3^*$  and  $\psi_4^*$ . Of these,  $\psi_3^*$  has two nodes, and  $\psi_4^*$ , the highest-energy molecular orbital, has three nodes between nuclei. Note that the number of nodes increases as the energy level of the orbital increases.

Quantum mechanical calculations show that the sum of energy levels of the two bonding butadiene molecular orbitals is slightly lower than the sum of two isolated alkene molecular orbitals. In other words, placing the four electrons in the two bonding molecular orbitals results in a more stable arrangement than placing them in two isolated alkene orbitals (Figure 14.3).



**Figure 14.2** Pi molecular orbitals in 1,3-butadiene. The asterisk on  $\psi_3^*$  and  $\psi_4^*$  indicates antibonding orbitals.



**Figure 14.3** An orbital view of 1,3-butadiene, showing the favorable pi bonding interaction between C2 and C3.

This “extra” stability is a consequence of a favorable bonding interaction across the C2–C3 bond of 1,3-butadiene present in  $\psi_1$ .

In describing the 1,3-butadiene molecular orbitals, we say that the pi electrons are *delocalized* over the entire pi framework rather than localized between two specific nuclei. Electron delocalization always leads to lower-energy orbitals and greater stability of the molecule.

## 14.4 Bond Lengths in 1,3-Butadiene

Further evidence for the special nature of conjugated dienes comes from data on bond lengths. Measurements show that the C2–C3 single bond in 1,3-butadiene has a length of 1.48 Å and that the two equivalent carbon–carbon double bonds have lengths of 1.34 Å (Table 14.2).

TABLE 14.2 Some Carbon–Carbon Bond Lengths

Bond	Bond length (Å)	Bond hybridization
CH <sub>3</sub> –CH <sub>3</sub>	1.54	C <sub>sp<sup>3</sup></sub> –C <sub>sp<sup>3</sup></sub>
H <sub>2</sub> C=CH <sub>2</sub>	1.33	C <sub>sp<sup>2</sup></sub> –C <sub>sp<sup>2</sup></sub>
H <sub>2</sub> C=CH–CH <sub>3</sub>	1.49	C <sub>sp<sup>2</sup></sub> –C <sub>sp<sup>3</sup></sub>
H <sub>2</sub> C=CH–CH=CH <sub>2</sub>	1.48	C <sub>sp<sup>2</sup></sub> –C <sub>sp<sup>2</sup></sub>
H <sub>2</sub> C=CHCH=CH <sub>2</sub>	1.34	C <sub>sp<sup>2</sup></sub> –C <sub>sp<sup>2</sup></sub>

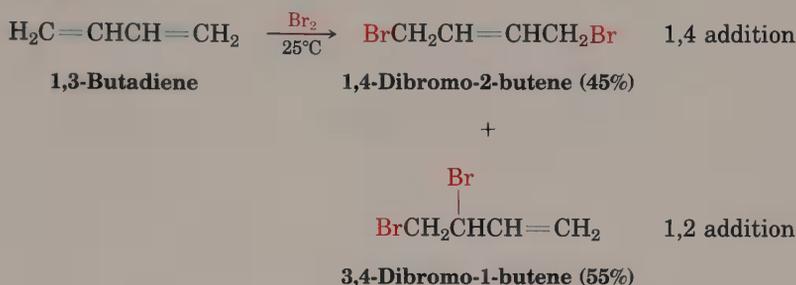
If we compare the length of the carbon–carbon single bond of 1,3-butadiene (1.48 Å) with that of ethane (1.54 Å), we find that the 1,3-butadiene single bond is shorter by 0.06 Å. Two explanations have been advanced to account for this bond shortening. One explanation ascribes the shortening of the 1,3-butadiene single bond to the delocalization of pi electrons in the bonding molecular orbitals. According to this view, pi orbital overlap across the C2–C3 bond results in partial double-bond character (Figure 14.3) and consequent bond shortening to a value midway between a pure single bond (1.54 Å) and a pure double bond (1.33 Å). The partial double-bond character of the C2–C3 bond is sufficient to stabilize the molecule but not to prevent bond rotation from occurring.

Alternatively, it can be argued that the shortened 1,3-butadiene single bond is a natural consequence of the orbital hybridization involved. The C2–C3 bond results from sigma overlap of two carbon  $sp^2$  orbitals, whereas a normal alkane bond results from overlap of two carbon  $sp^3$  orbitals. Since overlap of  $sp^2$  orbitals results in a single bond that has more *s* character than usual, the 1,3-butadiene single bond is a bit shorter and stronger than usual. Both explanations are valid, and both effects contribute to the bond shortening observed for 1,3-butadiene.

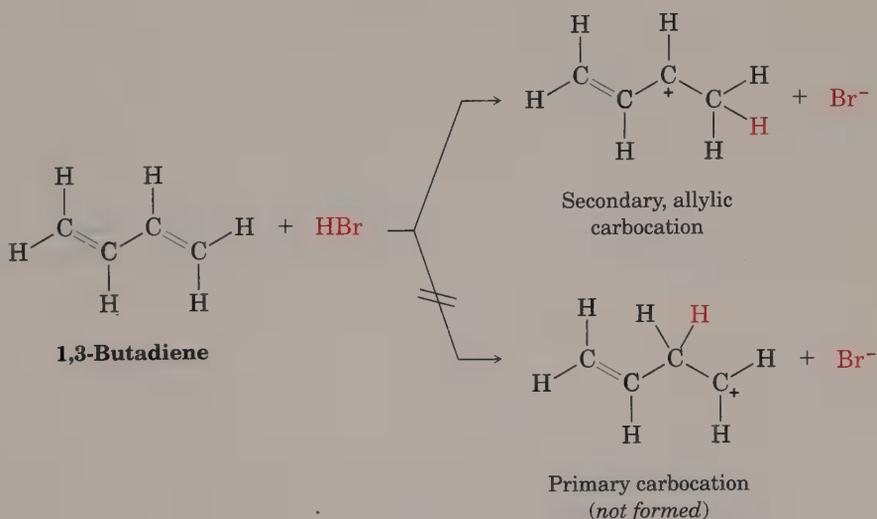


3-Bromo-1-butene is the normal Markovnikov product of **1,2-addition**, but 1-bromo-2-butene appears unusual. The double bond in this product has moved to a position between carbons 2 and 3, and HBr has added to carbons 1 and 4 (**1,4-addition**).

Many other electrophiles besides HBr add to conjugated dienes, and mixtures of products are formed in most cases. For example, Br<sub>2</sub> adds to 1,3-butadiene to give a mixture of 1,4-dibromo-2-butene and 3,4-dibromo-1-butene.

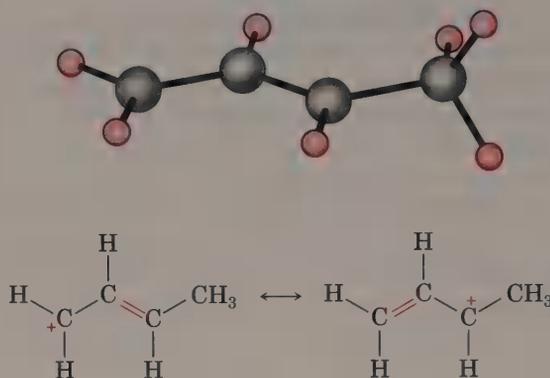


How can we account for the formation of the 1,4-addition products? The answer is that *allylic carbocations* are involved as intermediates in the reactions. When the electron-rich pi bond of 1,3-butadiene is protonated, two carbocation intermediates are possible: a primary carbocation and a secondary allylic cation (recall that *allylic* means “next to a double bond”). Since an allylic cation is stabilized by resonance (Section 11.9), it forms in preference to the less stable primary carbocation.



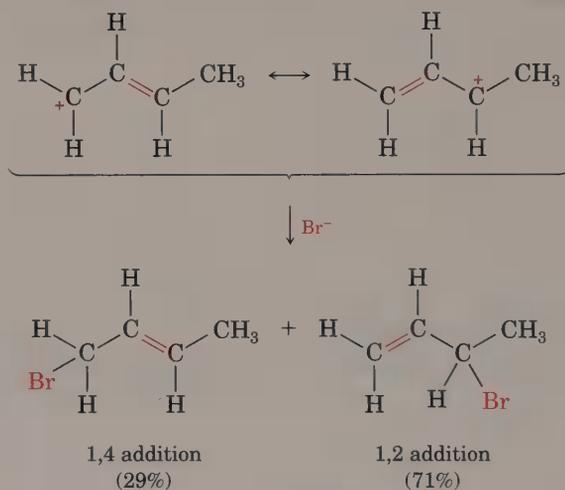
There are two alternative ways to account for the stability of allylic carbocations. Resonance theory (Section 10.7) offers a pictorial representation of the situation through the use of different resonance forms and says that the more resonance forms that are possible, the more stable the compound is. Thus, an allylic cation has two resonance forms and is more stable than a nonallylic cation. Neither of the two resonance forms is correct by

itself, of course; the true structure of the allylic cation is a combination, or resonance hybrid, of the two Kekulé structures.

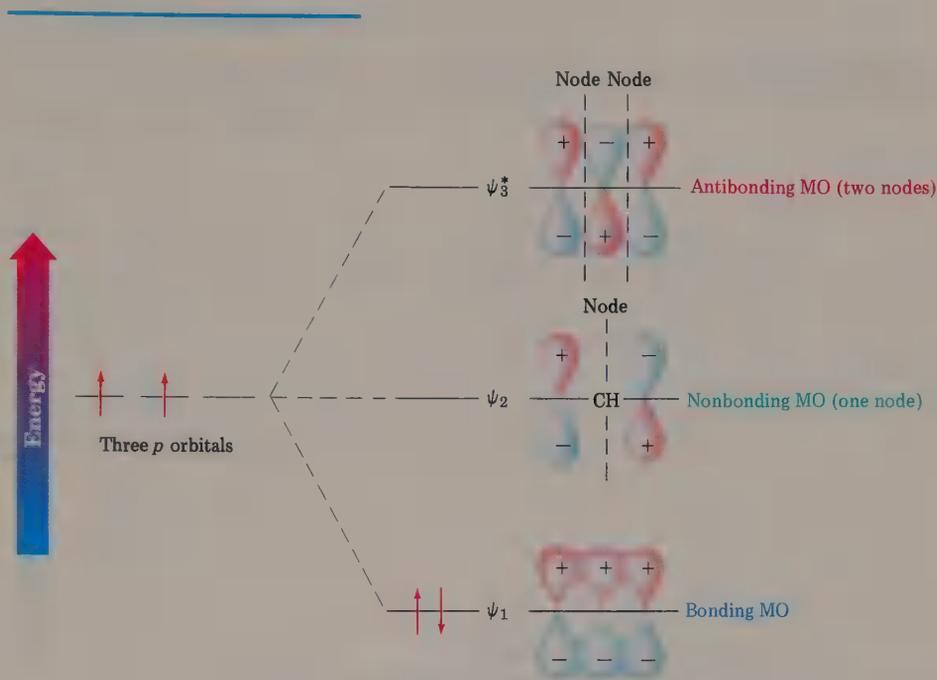


Two resonance forms  
of an allylic carbocation

When the allylic cation reacts with bromide ion to complete the electrophilic addition reaction, attack can occur either at carbon 1 or at carbon 3. Since both carbons share the positive charge, attack at both carbons occurs and a mixture of 1,2- and 1,4-addition products results.



An alternative way to account for the stability of allylic carbocations is to construct a molecular orbital description. When we allow three  $p$  orbitals to interact, three molecular orbitals are formed (Figure 14.4). One low-energy bonding orbital, one nonbonding orbital, and one high-energy antibonding orbital result. The two available pi electrons occupy the low-energy bonding orbital, indicating that a partial bond exists between carbons 2 and 3. Thus, the allylic cation is a conjugated system stabilized by electron delocalization in much the same way that 1,3-butadiene is stabilized by electron delocalization.



**Figure 14.4** Molecular orbitals of an allylic carbocation. The two available electrons occupy the low-energy bonding orbital, leading to a net stabilization.

It's important to realize that the two approaches just taken for describing allylic carbocations—the resonance approach and the molecular orbital approach—don't “compete” with each other. It's not that one approach is more “right” than the other; both are really just alternative ways of saying the same thing: Both are pictorial ways of visualizing a phenomenon that is best handled mathematically with wave functions. Sometimes one approach is more convenient, and sometimes the other is more convenient. We'll use both at various times.

**PROBLEM** .....

- 14.3 Give the structures of the likely products from reaction of one equivalent of HCl with 1,3-pentadiene. Show both 1,2 and 1,4 adducts.

**PROBLEM** .....

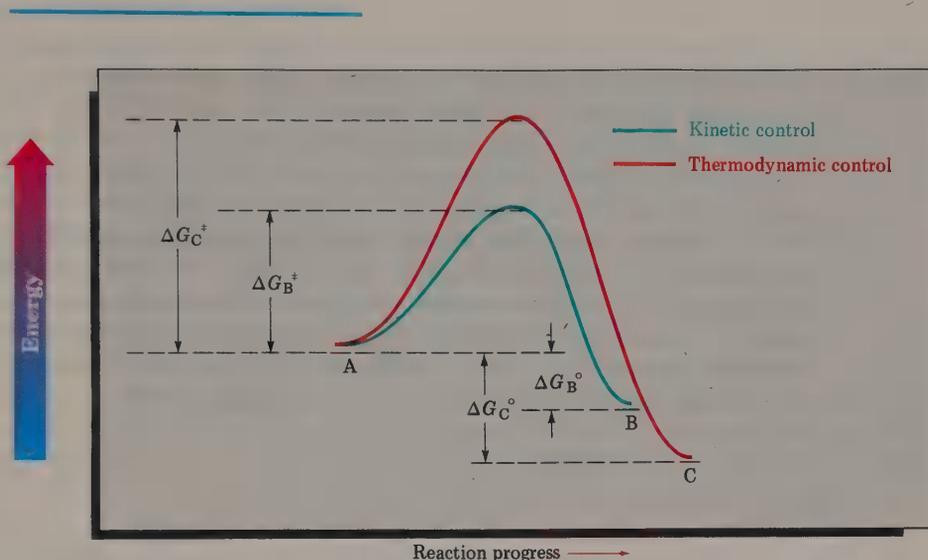
- 14.4 Look at the possible carbocation intermediates produced during addition of HCl to 1,3-pentadiene (Problem 14.3), and predict which of the 1,2 adducts predominates. Which 1,4 adduct predominates?

**PROBLEM** .....

- 14.5 Electrophilic addition of Br<sub>2</sub> to isoprene yields the following product mixture:

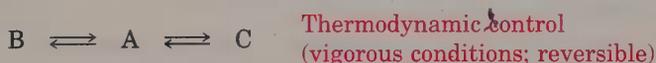


Let's assume that B forms faster than C (in other words, that  $\Delta G_{\text{B}}^{\ddagger} < \Delta G_{\text{C}}^{\ddagger}$ ) but that C is a more stable product than B (in other words, that  $\Delta G_{\text{C}}^{\circ} > \Delta G_{\text{B}}^{\circ}$ ). A reaction energy diagram for the two processes might look like that shown in Figure 14.5.



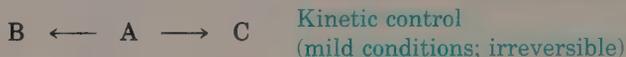
**Figure 14.5** A reaction energy diagram showing two competing reactions. The more stable product forms more slowly than the less stable product.

Let's first carry out the reaction under vigorous, high-temperature conditions so that both processes are reversible and an equilibrium is reached. Since C is more stable than B, C is the major product formed under the equilibrium reaction conditions. It doesn't matter that C forms more slowly than B, because the reaction conditions have been chosen such that B and C are formed reversibly, and therefore interconvert rapidly. *All that matters in a reversible reaction is the thermodynamic stability of the products at equilibrium.* Such reactions are said to be under **thermodynamic control**, or equilibrium control.

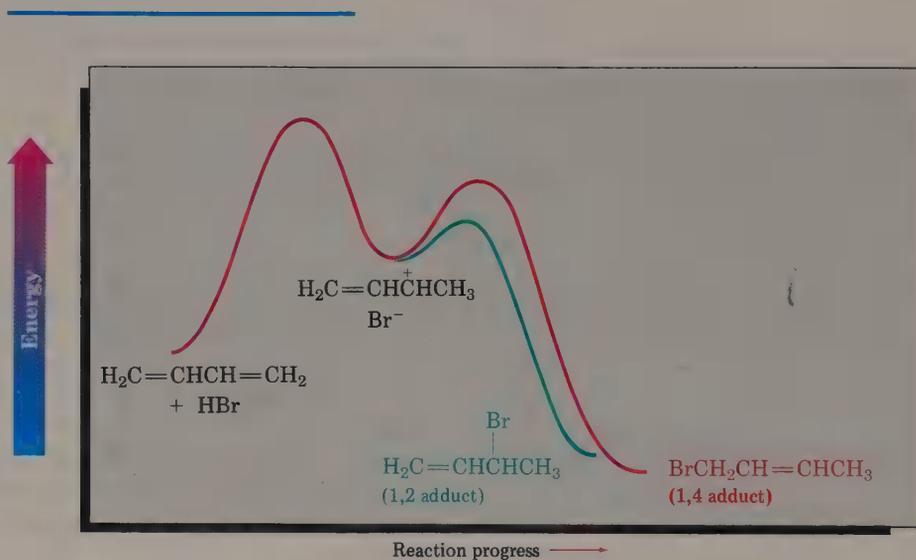


Now let's carry out the same reaction under milder, low-temperature conditions so that both processes are *irreversible* and no equilibrium is reached. In other words, there's enough energy present for the reactant A molecules to climb the energy barrier separating A from B, but not enough energy for product B and C molecules to climb the hill in the reverse direction and go back to reactant A. Since B forms faster than C, B is the major product. It doesn't matter that C is more stable than B, because the reaction

conditions have been chosen so that B and C are formed irreversibly and don't interconvert. *All that matters in an irreversible process is the reaction rate.* Such reactions are said to be under **kinetic control**.



We can now explain the effect of temperature on electrophilic addition reactions of conjugated dienes. Under mild, low-temperature conditions (0°C), HBr adds to 1,3-butadiene under kinetic control to give a 71:29 mixture of products with the 1,2 adduct predominating. Since these mild conditions don't allow the reaction to reach equilibrium, the product that forms faster predominates. Under more vigorous, high-temperature conditions (40°C), however, the reaction occurs reversibly under thermodynamic control to give a 15:85 mixture of products, with the more stable 1,4 adduct predominating. The higher temperature makes the second step of the addition process—reaction of the allylic carbocation intermediate with bromide ion—reversible, and an equilibrium mixture of products therefore results. Figure 14.6 shows the situation on a reaction energy diagram.



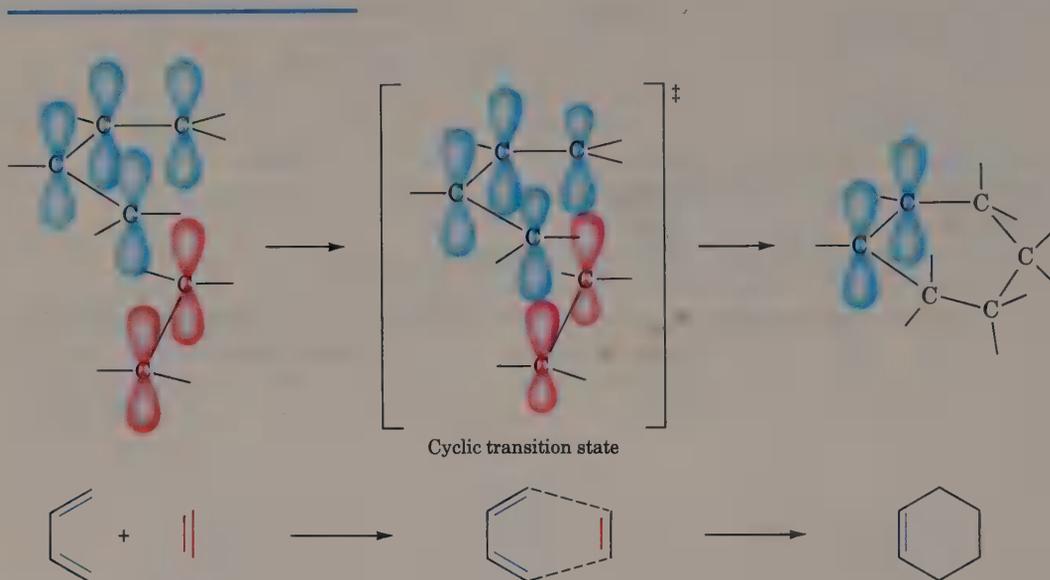
**Figure 14.6** Reaction energy diagram for the electrophilic addition of HBr to 1,3-butadiene. The 1,2 adduct is the kinetic product, and the 1,4 adduct is the thermodynamic product.

The electrophilic addition of HBr to 1,3-butadiene is an excellent example of how a change in experimental conditions can cause a change in the product of a reaction. The concept of thermodynamic control versus kinetic control is a valuable one that we can often use to advantage in the laboratory.



two reactants simply add together through a cyclic transition state in which both of the new carbon–carbon bonds form at the same time.

We can picture a Diels–Alder addition as occurring by head-on ( $\sigma$ ) overlap of the two isolated-alkene  $p$  orbitals with the two  $p$  orbitals on carbons 1 and 4 of the diene (Figure 14.7). This is, of course, a cyclic orientation of the reactants.



**Figure 14.7** Mechanism of the Diels–Alder cycloaddition reaction. The reaction occurs in a single step through a cyclic transition state in which the two new carbon–carbon bonds begin to form at the same time.

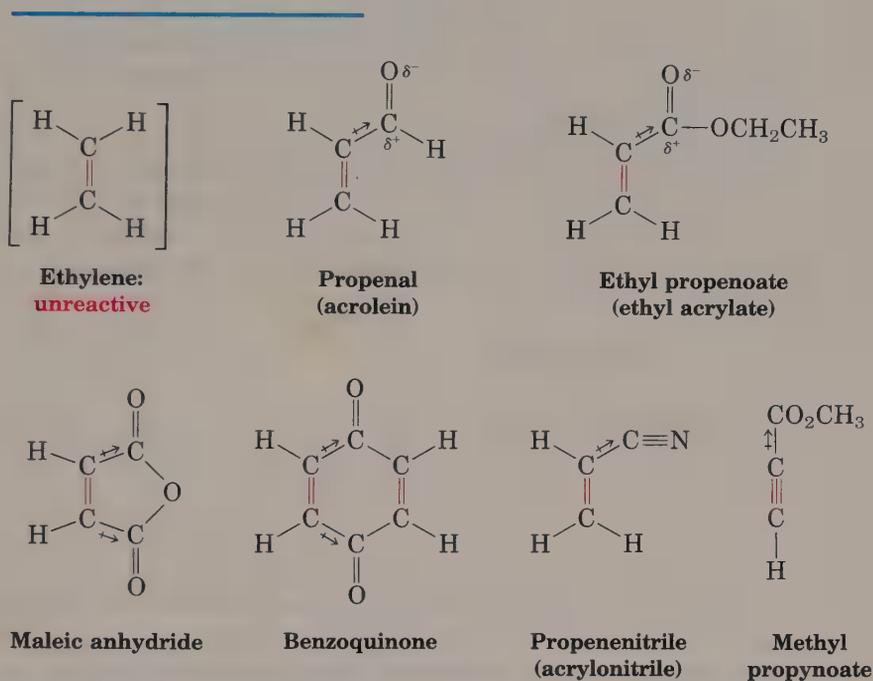
In the Diels–Alder transition state, the alkene carbons and carbons 1 and 4 of the diene rehybridize from  $sp^2$  to  $sp^3$  to form the two new single bonds. Carbons 2 and 3 of the diene remain  $sp^2$  hybridized to form the new double bond in the cyclohexene product. We'll study this mechanism at greater length in Chapter 30 but will concentrate for the present on learning more about the chemistry of the Diels–Alder reaction.

## 14.8 Characteristics of the Diels–Alder Reaction

### THE DIENOPHILE

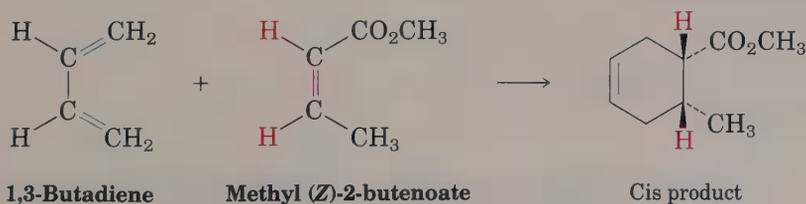
The Diels–Alder cycloaddition reaction takes place most rapidly and in highest yield if the alkene component, or **dienophile** (“diene lover”), is substituted by an electron-withdrawing group. Thus, ethylene itself is slug-

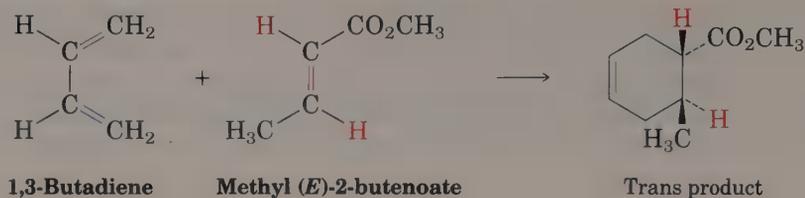
gish in the Diels–Alder reaction, but propenal, ethyl propenoate, maleic anhydride, benzoquinone, propenenitrile, and others are highly reactive (Figure 14.8). Note also that alkynes such as methyl propynoate can act as Diels–Alder dienophiles. In all of these cases, the dienophile multiple bond is next to the positively polarized carbon of a substituent that withdraws electrons.



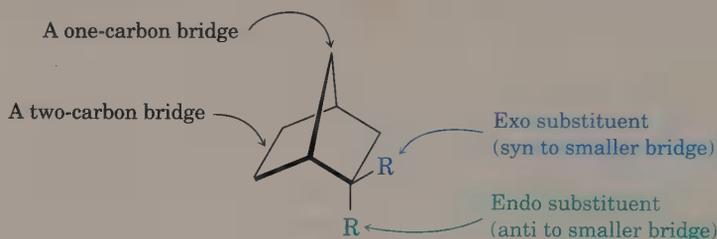
**Figure 14.8** Some Diels–Alder dienophiles. All contain electron-withdrawing substituents on the multiple bond.

One of the most important features of the Diels–Alder reaction is that it is *stereospecific*: The stereochemistry of the starting dienophile is maintained during the reaction, and a single product stereoisomer results. If we carry out the cycloaddition with a *cis* alkene such as methyl *cis*-2-butenate, only the *cis*-substituted cyclohexene product is formed. Conversely, Diels–Alder reaction with methyl *trans*-2-butenate yields only the *trans*-substituted cyclohexene product.

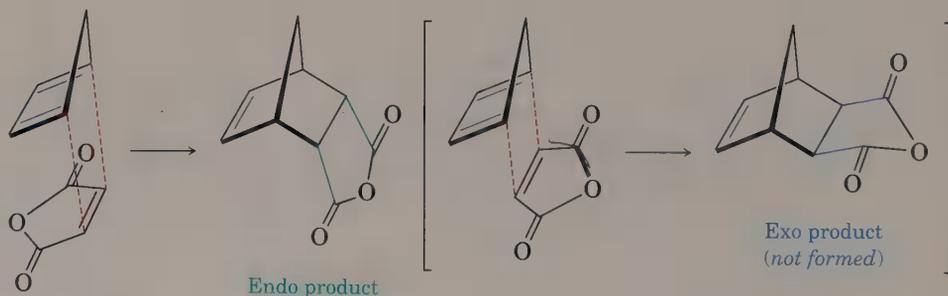




A second important stereochemical feature of the Diels–Alder reaction is that the diene and dienophile partners line up in such a way that the *endo* product rather than the alternative *exo* product is formed. The words *endo* and *exo* are used to indicate relative stereochemistry when referring to bicyclic structures like substituted norbornanes. A substituent on one bridge is said to be *exo* if it is *syn* (*cis*) to the smaller of the other two bridges and is said to be *endo* if it is *anti* (*trans*) to the smaller of the other two bridges. By analogy, a Diels–Alder reaction is said to have *endo* stereochemistry if the electron-withdrawing substituent on the dienophile ends up *syn* to the two-carbon unsaturated bridge in the product.



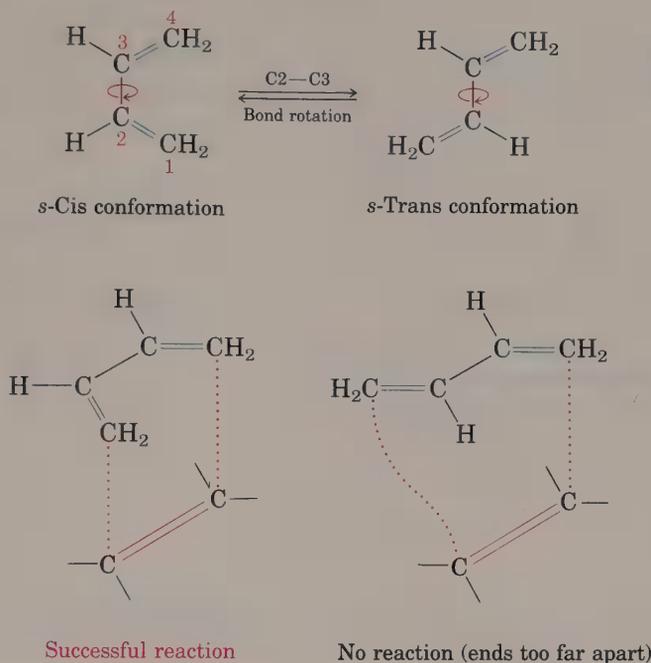
When diene and dienophile approach each other in a Diels–Alder reaction, the best orbital overlap is achieved when the reactants lie directly on top of one another, a geometry that leads to *endo* rather than *exo* product. In the reaction of 1,3-cyclopentadiene with maleic anhydride, for example, the following result is obtained:



## THE DIENE

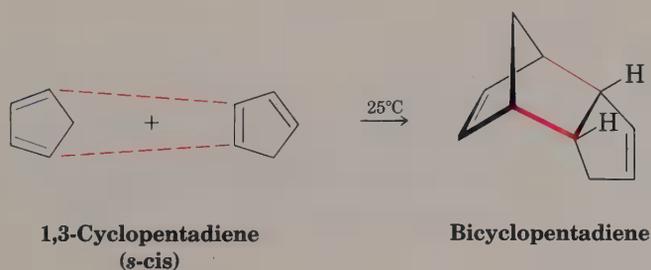
A diene must be able to adopt an *s-cis* conformation (a “*cis*-like” conformation about the single bond) in order to undergo the Diels–Alder reaction. Only

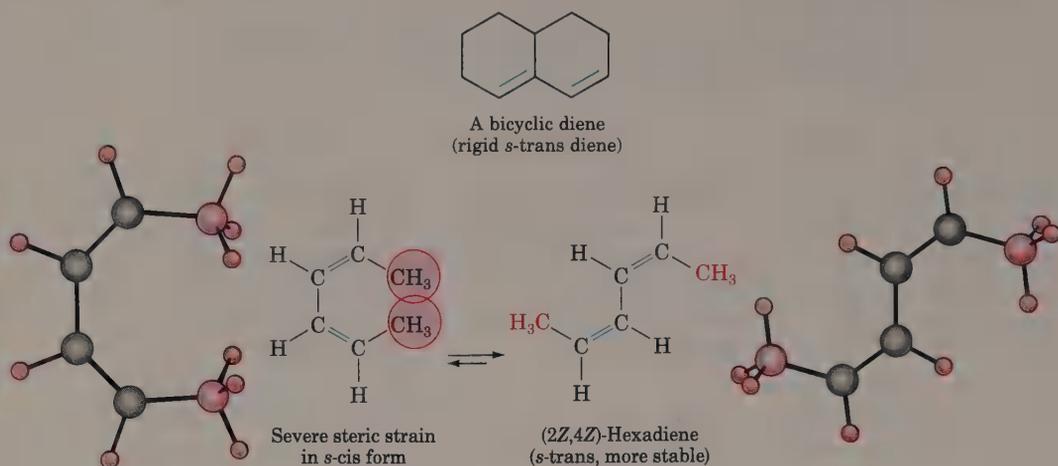
in the *s-cis* conformation are carbons 1 and 4 of the diene close enough to react through a cyclic transition state to give a new ring. In the alternative *s-trans* conformation, the ends of the diene partner are too far apart to overlap the dienophile *p* orbitals successfully.



Examples of some dienes that can't adopt *s-cis* conformation and therefore don't undergo Diels–Alder reaction are shown in Figure 14.9. In the case of the bicyclic diene, the double bonds are rigidly fixed in the *s-trans* arrangement by geometric constraints of the rings. In the case of (2*Z*,4*Z*)-hexadiene, severe steric strain between the two methyl groups prevents the molecule from adopting *s-cis* geometry.

In contrast to the unreactive *s-trans* dienes, certain other dienes are rigidly fixed in the correct *s-cis* geometry and are therefore highly reactive in the Diels–Alder cycloaddition reaction. Such, for example, is the case with cyclopentadiene. Cyclopentadiene is so reactive, in fact, that it reacts with itself! At room temperature, cyclopentadiene *dimerizes*: one molecule acts as diene and another acts as dienophile in a self Diels–Alder reaction.

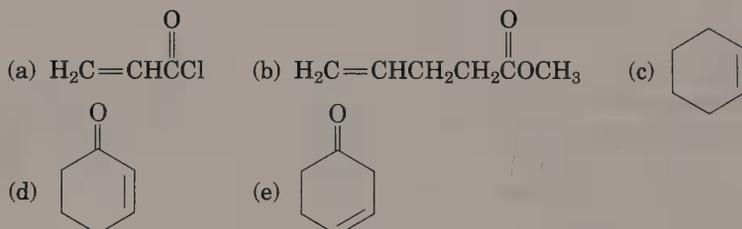




**Figure 14.9** Two *s-trans* dienes that can't undergo Diels–Alder reactions.

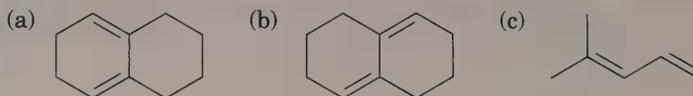
**PROBLEM** .....

**14.9** Which of these alkenes would you expect to be good Diels–Alder dienophiles?



**PROBLEM** .....

**14.10** Which of the following dienes have an *s-cis* conformation and which an *s-trans* conformation? Of the *s-trans* dienes, which can readily rotate to *s-cis*?



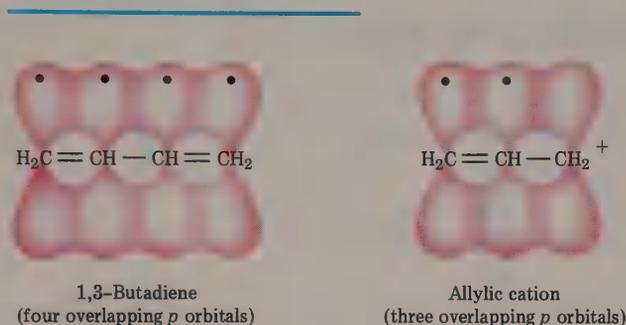
**PROBLEM** .....

**14.11** Although cyclopentadiene is highly reactive toward Diels–Alder cycloaddition reactions, 1,3-cyclohexadiene is less reactive, and 1,3-cycloheptadiene is nearly inert. Can you suggest a reason for this reactivity order? (Building molecular models should be helpful.)

.....

## 14.9 Other Conjugated Systems

A conjugated system was defined earlier in this chapter as one that consists of alternating double and single bonds. After considering a molecular orbital description of 1,3-butadiene, however, we might now more accurately describe a conjugated system as one that consists of *an extended series of overlapping p orbitals* (Figure 14.10). Thus, a 1,3-diene and an allylic cation are both examples of conjugated systems, but there are other kinds as well.



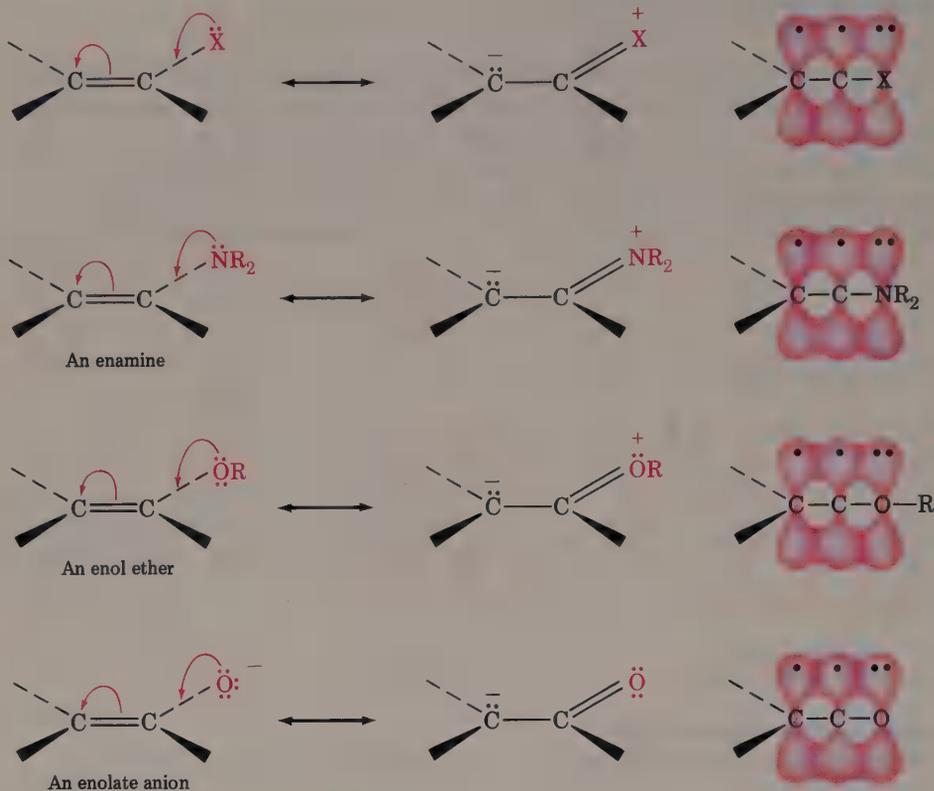
**Figure 14.10** Conjugated systems such as 1,3-butadiene and allylic cations contain an extended series of overlapping  $p$  orbitals.

Some of the most important kinds of conjugated systems result from overlap of double-bond  $p$  orbitals with a  $p$  orbital on a neighboring atom such as oxygen, nitrogen, or halogen. The system can be either negatively charged or neutral (Figure 14.11).

These conjugated systems will be examined in more detail in later chapters. For the present, it's sufficient to note that the electronic nature of the alkene  $\pi$  bond is greatly affected by conjugation with a filled neighboring orbital. As the resonance forms shown in Figure 14.11 indicate, conjugation of the double bond with a filled neighboring orbital greatly increases the electron density of the carbon-carbon double bond. Thus, enol ethers, enamines, and enolate ions are all much more strongly nucleophilic than normal isolated alkenes.



This resonance form puts extra electron density on *carbon*, making the carbon atom nucleophilic



**Figure 14.11** Some conjugated systems resulting from overlap of double-bond  $p$  orbitals with filled  $p$  orbitals on neighboring atoms.

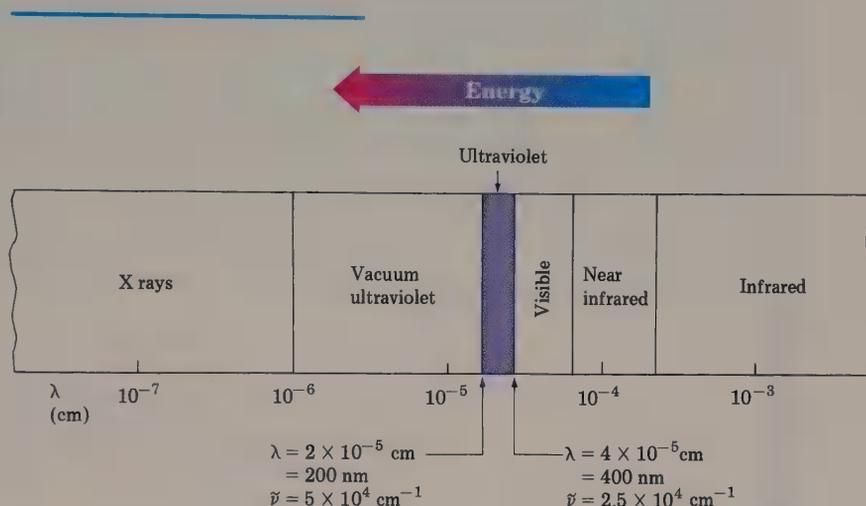
## 14.10 Structure Determination of Conjugated Systems: Ultraviolet Spectroscopy

Infrared, nuclear magnetic resonance, and mass spectroscopy all help in the structure determination of conjugated systems. In addition to these three generally useful spectroscopic techniques, there's a fourth—**ultraviolet (UV) spectroscopy**—that is applicable solely to conjugated systems.

- |  |  |
|--|--|
| 1. Infrared spectroscopy                   | Functional groups present                  |
| 2. Mass spectroscopy                       | Molecular size and formula                 |
| 3. Nuclear magnetic resonance spectroscopy | Carbon–hydrogen framework                  |
| 4. Ultraviolet spectroscopy                | Nature of conjugated $\pi$ electron system |

Ultraviolet spectroscopy is less commonly used than the other three spectroscopic techniques because of the rather specialized information it gives. We'll therefore study it only briefly.

The ultraviolet region of the electromagnetic spectrum extends from the low wavelength end of the visible region ( $4 \times 10^{-5}$  cm) down to  $10^{-6}$  cm, but the narrow range from  $2 \times 10^{-5}$  cm to  $4 \times 10^{-5}$  cm is the portion of greatest interest to organic chemists. Absorptions in this region are usually measured in nanometers, nm ( $1 \text{ nm} = 10^{-9} \text{ m} = 10^{-7} \text{ cm}$ ). Thus, the ultraviolet range of interest is from 200 to 400 nm (Figure 14.12).



**Figure 14.12** The ultraviolet (UV) region of the electromagnetic spectrum.

We saw in our discussion of infrared spectroscopy (Section 12.9) that when an organic molecule is irradiated with electromagnetic energy, the radiation either is absorbed by the compound or passes through, depending on the radiation's exact energy. When infrared radiation is used, the energy absorbed corresponds to the amount necessary to increase molecular bending and stretching vibrations in functional groups. When ultraviolet radiation is used, the energy absorbed by a molecule corresponds to the amount necessary to excite electrons from one molecular orbital to another. Let's see what this means by looking first at 1,3-butadiene.

**PROBLEM** .....

- 14.12** Calculate the energy range of electromagnetic radiation in the ultraviolet region of the spectrum from 200 to 400 nm wavelength. Recall

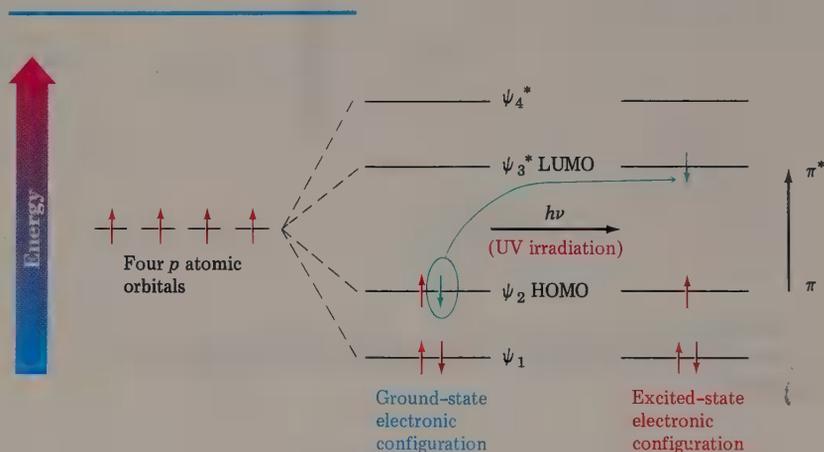
$$E = \frac{Nhc}{\lambda} = \frac{2.86 \times 10^{-3} \text{ kcal/mol}}{\lambda \text{ (cm)}}$$

## PROBLEM .....

- 14.13 How does the energy you calculated (Problem 14.12) for ultraviolet radiation compare with the values calculated previously for infrared spectroscopy and nuclear magnetic resonance spectroscopy?
- .....

## 14.11 Ultraviolet Spectrum of 1,3-Butadiene

1,3-Butadiene has four pi molecular orbitals. The two lower-energy bonding molecular orbitals are fully occupied in the ground state, and the two higher-energy antibonding molecular orbitals are unoccupied, as illustrated in Figure 14.13.

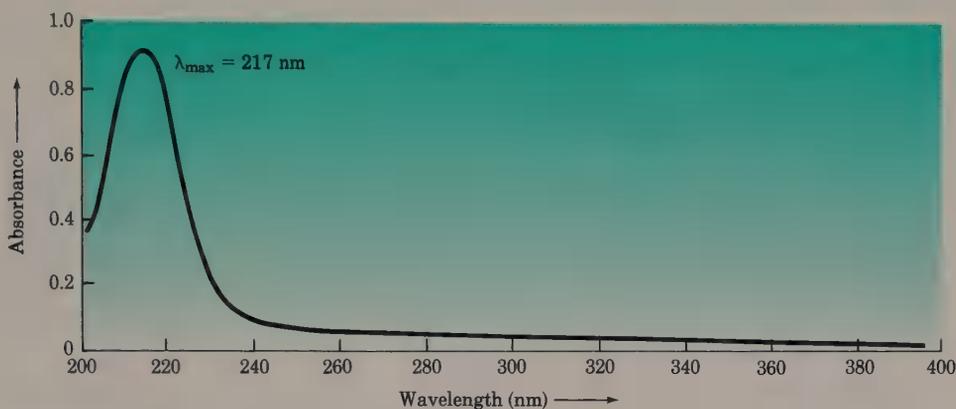


**Figure 14.13** Ultraviolet excitation of 1,3-butadiene results in the promotion of an electron from a bonding orbital ( $\psi_2$ ) to an antibonding orbital ( $\psi_3^*$ ).

On irradiation with ultraviolet light ( $h\nu$ ), 1,3-butadiene absorbs energy, and a pi electron is promoted from  $\psi_2$ , the highest occupied molecular orbital (HOMO), to  $\psi_3^*$ , the lowest unoccupied molecular orbital (LUMO). Since the electron is promoted from a bonding  $\pi$  molecular orbital to an antibonding  $\pi^*$  molecular orbital, we call this a  $\pi \rightarrow \pi^*$  excitation (read as “pi to pi star”). The energy gap between the HOMO and the LUMO of 1,3-butadiene is such that ultraviolet light of 217 nm wavelength is required to accomplish the  $\pi \rightarrow \pi^*$  electronic transition.

In practice, an ultraviolet spectrum is recorded by irradiating the sample with ultraviolet light of continuously changing wavelength. When the wavelength of light corresponds to the energy level required to excite an electron to a higher level, energy is absorbed. This absorption is detected

and displayed on a chart that plots wavelength versus percent radiation absorbed (Figure 14.14). Note that UV spectra differ from IR spectra in the way they are recorded. IR spectra are usually displayed so that the “baseline” corresponding to zero absorption runs across the top of the chart and a valley indicates an absorption. UV spectra are displayed with the baseline at the bottom of the chart so that a peak indicates an absorption.



**Figure 14.14** The ultraviolet spectrum of 1,3-butadiene,  $\lambda_{\max} = 217$  nm.

The exact amount of UV light absorbed is expressed as the sample’s **molar absorptivity**, or **extinction coefficient**,  $\epsilon$ , defined by the equation

$$\text{Molar absorptivity } \epsilon = \frac{A}{C \times l}$$

where  $A$  = Absorbance, expressed as  $\log\left(\frac{I_0}{I}\right)$  where  $I_0$  is the intensity of the incident light and  $I$  is the intensity of the light transmitted through the sample

$C$  = Concentration in mol/L

$l$  = Sample path length in cm

Molar absorptivity is a physical constant, characteristic of the particular substance being observed and thus characteristic of the particular pi electron system present in the sample. Typical values for conjugated dienes are in the range  $\epsilon = 10,000$ – $25,000$ .

Unlike infrared spectra and nuclear magnetic resonance spectra, which show many absorption lines for a given molecule, ultraviolet spectra are usually quite simple—often only a single peak is seen. The peak is usually broad, however, and we identify its position by noting the wavelength ( $\lambda$ ) at the very top of the peak ( $\lambda_{\max}$ ; read as “lambda max”).

## PROBLEM .....

- 14.14 Knowledge of molar absorptivities is particularly important in biochemistry where UV spectroscopy can provide an extremely sensitive method of analysis. For example, imagine that you wanted to determine the concentration of vitamin A in a sample. If pure vitamin A has  $\lambda_{\text{max}} = 325$  ( $\epsilon = 50,100$ ) in a cell with a path length of 1.00 cm, what is the vitamin A concentration in a sample whose absorbance at 325 nm is  $A = 0.735$ ?
- .....

## 14.12 Interpreting Ultraviolet Spectra: The Effect of Conjugation

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The exact wavelength of radiation necessary to effect the  $\pi \rightarrow \pi^*$  transition in a conjugated molecule depends on the energy gap between molecular orbitals (HOMO and LUMO), which in turn depends on the exact nature of the conjugated system. Thus, by measuring the ultraviolet spectrum of an unknown, we can derive structural information about the nature of any conjugated pi electron system present in a sample.

One of the most important factors affecting the wavelength of ultraviolet absorption by a given molecule is the extent of conjugation. Molecular orbital calculations show that the energy difference between HOMO and LUMO decreases as the extent of conjugation increases; thus, 1,3-butadiene shows an absorption at  $\lambda_{\text{max}} = 217$  nm, 1,3,5-hexatriene absorbs at  $\lambda_{\text{max}} = 258$  nm, and 1,3,5,7-octatetraene has  $\lambda_{\text{max}} = 290$  nm. [*Remember*: Longer wavelength means lower energy.]

Other kinds of conjugated systems besides dienes and polyenes also show ultraviolet absorptions. For example, conjugated enones and aromatic rings exhibit characteristic ultraviolet absorptions that aid in structure determination. More will be said about such compounds later when the specific functional groups are discussed in detail. The ultraviolet absorption maxima of some representative conjugated molecules are given in Table 14.3.

In addition to the  $\pi \rightarrow \pi^*$  absorptions just discussed, other electronic transitions are also observed in ultraviolet spectroscopy. Compounds with nonbonding electrons, such as the lone-pair electrons on oxygen, nitrogen, and halogen, also show weak ultraviolet absorption. In these cases, a nonbonding electron ( $n$ ) is promoted to an antibonding orbital ( $\pi^*$ ). In acetone, for example, a nonbonding lone-pair electron on oxygen is excited by ultraviolet irradiation into the carbonyl antibonding  $\pi^*$  orbital. The resultant  $n \rightarrow \pi^*$  transition shows an absorption peak at  $\lambda_{\text{max}} = 272$  nm but is quite weak compared to the usual  $\pi \rightarrow \pi^*$  absorption seen in conjugated systems.

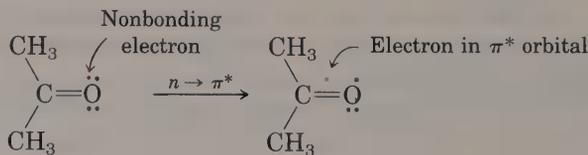
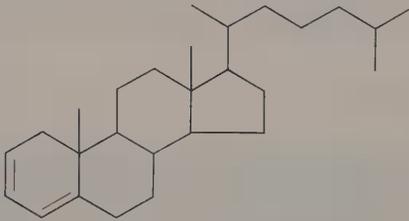


TABLE 14.3 Ultraviolet Absorption Maxima of Some Conjugated Molecules

Name	Structure	$\lambda_{\max}$ (nm)
Ethylene	$\text{H}_2\text{C}=\text{CH}_2$	171
Cyclohexene		182
2-Methyl-1,3-butadiene	$\begin{array}{c} \text{CH}_3 \\   \\ \text{H}_2\text{C}=\text{C}-\text{CH}=\text{CH}_2 \end{array}$	220
1,3-Cyclohexadiene		256
1,3,5-Hexatriene	$\text{H}_2\text{C}=\text{CH}-\text{CH}=\text{CH}-\text{CH}=\text{CH}_2$	258
1,3,5,7-Octatetraene	$\text{H}_2\text{C}=\text{CH}-\text{CH}=\text{CH}-\text{CH}=\text{CH}-\text{CH}=\text{CH}_2$	290
2,4-Cholestadiene		275
3-Buten-2-one	$\begin{array}{c} \text{CH}_3 \\   \\ \text{H}_2\text{C}=\text{CH}-\text{C}=\text{O} \end{array}$	219
Benzene		254
Naphthalene		275

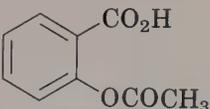
## PROBLEM .....

14.15 Which of the following compounds would you expect to show ultraviolet absorptions in the 200–400 nm range?

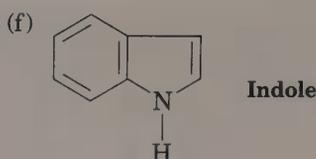
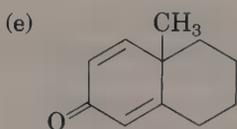
(a) 1,4-Cyclohexadiene

(b) 1,3-Cyclohexadiene

(c)  $\text{H}_2\text{C}=\text{CH}-\text{C}\equiv\text{N}$

(d) 

Aspirin



### 14.13 Colored Organic Compounds and the Chemistry of Vision

Why are some organic compounds colored but others aren't? Why is  $\beta$ -carotene (from carrots) orange but benzene is colorless? The answers have to do both with the structures of colored molecules and with the way we perceive light.

The visible region of the electromagnetic spectrum is adjacent to the ultraviolet region, extending from approximately 400 to 800 nm. Colored compounds like  $\beta$ -carotene have such extended systems of conjugation that

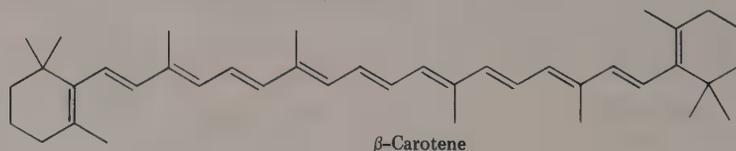
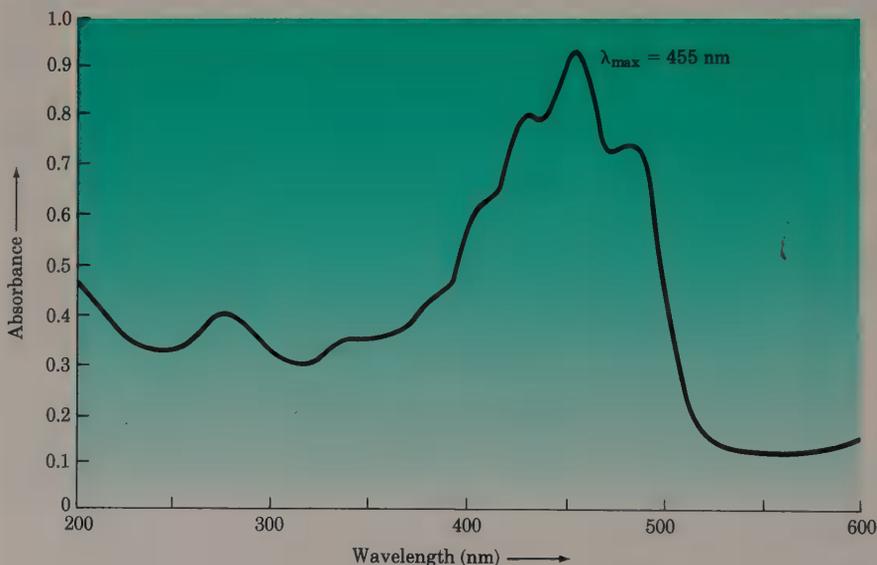


Figure 14.15 Ultraviolet spectrum of  $\beta$ -carotene.

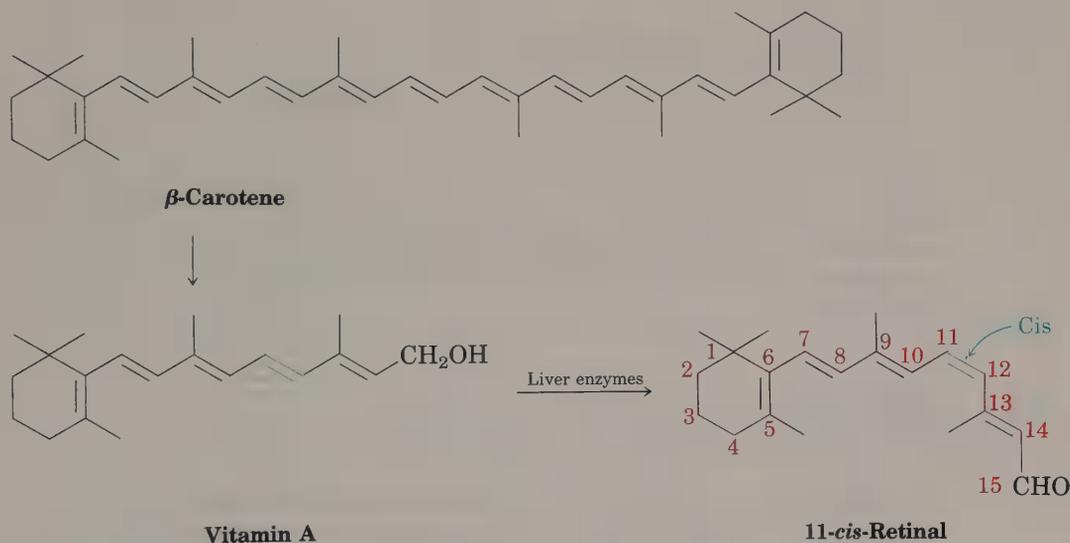
their “UV” absorptions actually extend out into the visible region.  $\beta$ -Carotene, for example, has 11 double bonds in conjugation, and its absorption occurs at  $\lambda_{\text{max}} = 455 \text{ nm}$  (Figure 14.15).

Ordinary “white” light, from the sun or from a lamp, consists of all wavelengths in the visible region. When white light strikes  $\beta$ -carotene, the wavelengths from 400 to 500 nm (blue) are absorbed, while all other wavelengths are transmitted and can reach our eyes. We therefore see the white light with the blue subtracted out, and we perceive a yellow-orange color for  $\beta$ -carotene. (The yellow-orange coloration accounts for the use of  $\beta$ -carotene as a food-coloring agent in margarine.)

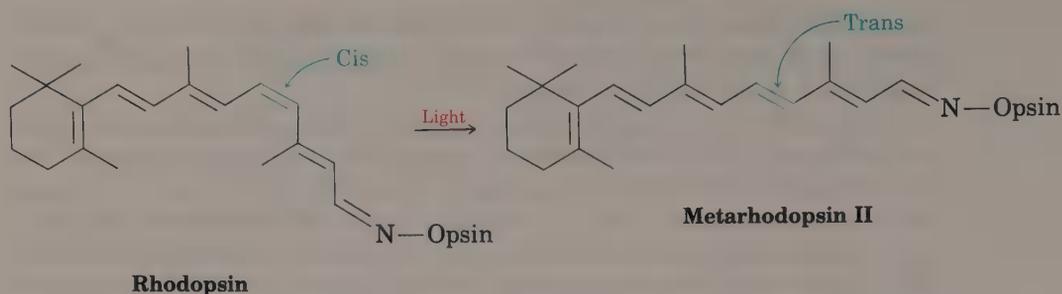
What is true for  $\beta$ -carotene is true for all other colored organic compounds: All have an extended system of pi electron conjugation that gives rise to an absorption in the visible region of the electromagnetic spectrum.

## THE CHEMISTRY OF VISION

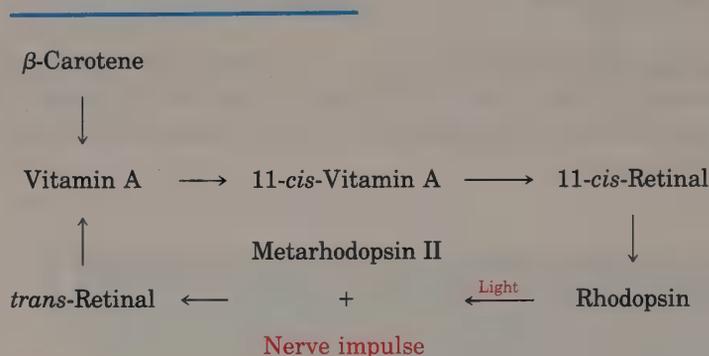
Not only are conjugated molecules colored, they also make up the light-sensitive substances on which the visual systems of all living things are based. 11-*cis*-Retinal, synthesized in the liver from dietary vitamin A, is the key substance.



There are two types of light-sensitive receptor cells in the eye, *rod* cells and *cone* cells. Rod cells are primarily responsible for seeing in dim light, whereas cone cells are responsible for seeing in bright light and for the perception of bright colors. In the rod cells of the eye, 11-*cis*-retinal is converted into *rhodopsin*, a light-sensitive substance formed from the protein *opsin* and 11-*cis*-retinal. When light strikes the rod cell, isomerization of the C11–C12 double bond occurs and *trans*-rhodopsin, called metarhodopsin II, is produced. This *cis*–*trans* isomerization of rhodopsin is accompanied by a change in molecular geometry, which in turn causes a nerve impulse to be sent to the brain where it is perceived as vision.



Metarhodopsin II is then recycled back into rhodopsin by a multistep sequence involving cleavage into all-*trans*-retinal, conversion to vitamin A, *cis*-*trans* isomerization to 11-*cis*-vitamin A, and conversion back to 11-*cis*-retinal (Figure 14.16).



**Figure 14.16** The visual cycle. The actual series of events is more complicated than the diagram indicates, involving several intermediate steps in the light-induced conversion of rhodopsin into metarhodopsin II.

## 14.14 Summary and Key Words

A **conjugated diene** is one that contains alternating double and single bonds. Thus, 1,3-butadiene is conjugated, whereas 1,4-pentadiene is non-conjugated.

One important difference between conjugated and nonconjugated dienes is that conjugated dienes are somewhat more stable than we might expect. This unexpected stability can be explained by a molecular orbital description in which four *p* atomic orbitals overlap to form four **molecular orbitals**, denoted  $\psi_1$ ,  $\psi_2$ ,  $\psi_3^*$ , and  $\psi_4^*$ . Two of the molecular orbitals are **bonding** and two are **antibonding**. When four electrons are added, only the two bonding orbitals,  $\psi_1$  and  $\psi_2$ , are filled. The two antibonding orbitals are unoccupied.

Calculations show that the sum of the energy levels of the two bonding 1,3-butadiene molecular orbitals is lower (that is, the system is more stable)

than the sum of the energy levels of two isolated double bonds. The reason for this extra stability is a bonding interaction between carbons 2 and 3, which introduces some partial double-bond character.

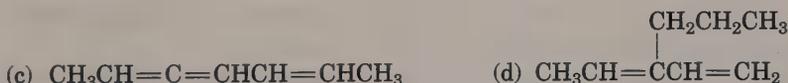
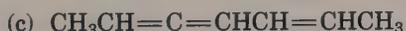
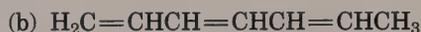
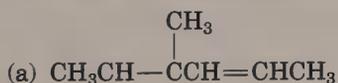
Conjugated dienes undergo two reactions not observed for nonconjugated dienes. The first of these is **1,4 addition** of electrophiles. When 1,3-butadiene is treated with an electrophile like HCl, **1,2** and **1,4 adducts** are formed. Both products are formed from the same resonance-stabilized allylic carbocation intermediate and are produced in varying ratios depending on the reaction conditions. The 1,2 adduct is usually formed faster and predominates at low temperature (**kinetic control**), whereas the 1,4 adduct usually predominates when equilibrium is achieved at higher temperatures (**thermodynamic control**).

The second reaction unique to conjugated dienes is **Diels–Alder cycloaddition**. Conjugated dienes react with electron-poor alkenes (dienophiles) in a single step through a cyclic transition state to yield a cyclohexene product. This is an example of the general class of processes called **pericyclic reactions**, which have neither polar nor radical mechanisms. Diels–Alder reactions can occur only if the diene is able to adopt an **s-cis conformation**.

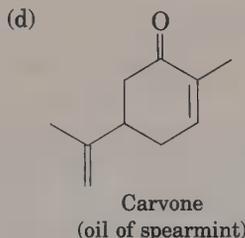
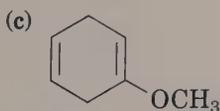
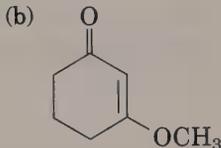
**Ultraviolet (UV) spectroscopy** is a method of structure determination applicable only to conjugated systems. When a conjugated molecule is irradiated with ultraviolet light, energy absorption occurs and a pi electron is promoted from the highest occupied molecular orbital (**HOMO**) to the lowest unoccupied molecular orbital (**LUMO**). For 1,3-butadiene, radiation of  $\lambda_{\text{max}} = 217 \text{ nm}$  is required. As a general rule, the greater the extent of conjugation, the less the energy needed (longer wavelength radiation).

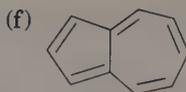
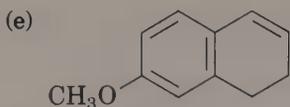
## ADDITIONAL PROBLEMS .....

14.16 Provide IUPAC names for the following alkenes:



14.17 Circle any conjugated portions of these molecules:





14.18 What product(s) would you expect to obtain from reaction of 1,3-cyclohexadiene with each of the following?

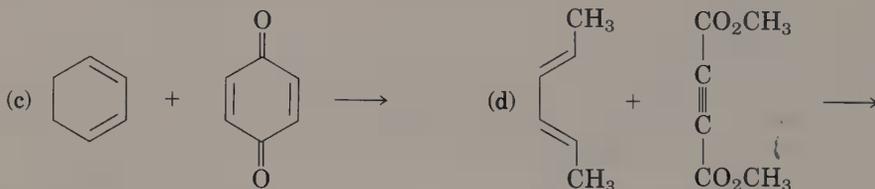
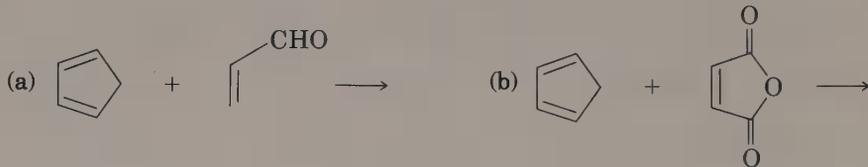
- (a) 1 mol  $\text{Br}_2$  in  $\text{CCl}_4$  (b)  $\text{O}_3$  followed by Zn  
 (c) 1 mol HCl in ether (d) 1 mol DCl in ether  
 (e) 3-Buten-2-one ( $\text{H}_2\text{C}=\text{CHCOCH}_3$ ) (f) Excess  $\text{OsO}_4$ , followed by  $\text{NaHSO}_3$

14.19 Draw and name the six possible diene isomers of formula  $\text{C}_5\text{H}_8$ . Which of the six are conjugated dienes?

14.20 Treatment of 3,4-dibromohexane with strong base leads to loss of two equivalents of HBr and formation of a product with formula  $\text{C}_6\text{H}_{10}$ . Three products are possible. Name each of the three, and tell how you would use  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectroscopy to help you identify the product. How would you use ultraviolet spectroscopy?

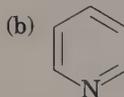
14.21 Would you expect allene,  $\text{H}_2\text{C}=\text{C}=\text{CH}_2$ , to show a UV absorption in the 200–400 nm range? Explain.

14.22 Predict the products of these Diels–Alder reactions:

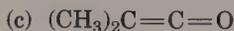


14.23 How do you account for the fact that *cis*-1,3-pentadiene is much less reactive than *trans*-1,3-pentadiene in the Diels–Alder reaction?

14.24 Which of the following compounds would you expect to have  $\pi \rightarrow \pi^*$  ultraviolet absorptions in the 200–400 nm range?



Pyridine

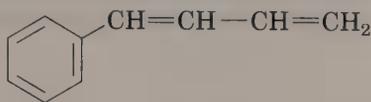


A ketene

14.25 Would you expect a conjugated diyne such as 1,3-butadiyne to undergo Diels–Alder reaction with a dienophile? Explain.

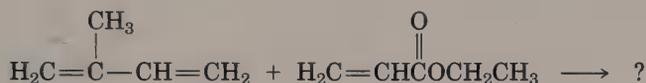
14.26 Propose a structure for a conjugated diene that gives the same product from both 1,2 and 1,4 addition of HBr.

- 14.27 Draw the products resulting from addition of 1 mol HCl to 1-phenyl-1,3-butadiene. Which product or products would you expect to predominate, and why?



1-Phenyl-1,3-butadiene

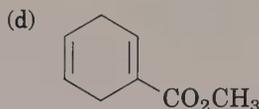
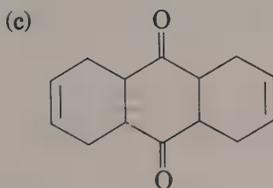
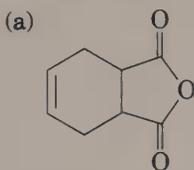
- 14.28 Reaction of isoprene (2-methyl-1,3-butadiene) with ethyl propenoate gives a mixture of two Diels–Alder adducts. Show the structure of each, and explain why a mixture is formed.



- 14.29 Rank the following dienophiles in order of their expected reactivity in the Diels–Alder reaction. Explain.



- 14.30 How would you use Diels–Alder reactions to prepare these products? Show the starting dienes and dienophiles in each case.



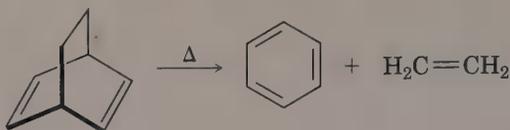
- 14.31 Aldrin, a chlorinated insecticide now banned by the U.S. Environmental Protection Agency, can be made by Diels–Alder reaction of hexachloro-1,3-cyclopentadiene with norbornadiene. Formulate the reaction.



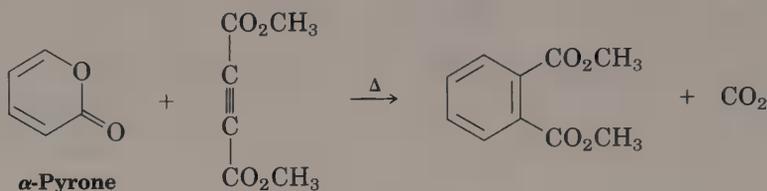
Norbornadiene

- 14.32 Norbornadiene (Problem 14.31) can be prepared by reaction of chloroethylene with cyclopentadiene, followed by treatment of the product with sodium ethoxide. Formulate the overall scheme, and identify the two kinds of reactions.

- 14.33 We've seen that the Diels–Alder cycloaddition reaction is a pericyclic process that occurs in a concerted manner through a cyclic transition state. Depending on the exact energy levels of products and reactants, the Diels–Alder reaction can sometimes be made reversible. How can you account for the following reaction?



- 14.34 Propose a mechanism to explain the following reaction. [Hint: Consider the reversibility of the Diels–Alder reaction, Problem 14.33.]

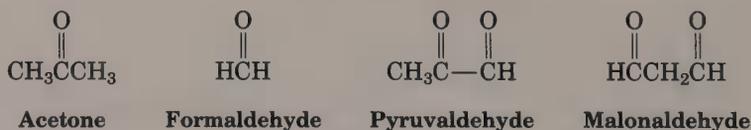


- 14.35 The following ultraviolet absorption maxima have been measured:

	$\lambda_{\max}$ (nm)
1,3-Butadiene	217
2-Methyl-1,3-butadiene	220
1,3-Pentadiene	223
2,3-Dimethyl-1,3-butadiene	226
2,4-Hexadiene	227
2,4-Dimethyl-1,3-pentadiene	232
2,5-Dimethyl-2,4-hexadiene	240

What conclusion can you draw from these data concerning the effect of alkyl substitution on ultraviolet absorption maxima? Approximately what effect does each added alkyl group have?

- 14.36 1,3,5-Hexatriene has  $\lambda_{\max} = 258$  nm. In light of your answer to Problem 14.35, approximately where would you expect 2,3-dimethyl-1,3,5-hexatriene to absorb? Explain.
- 14.37  $\beta$ -Ocimene is a pleasant-smelling hydrocarbon found in the leaves of certain herbs. It has the molecular formula  $C_{10}H_{16}$  and exhibits an ultraviolet absorption maximum at 232 nm. On catalytic hydrogenation over palladium, 2,6-dimethyloctane is obtained. Ozonolysis of  $\beta$ -ocimene, followed by treatment with zinc and acetic acid, produces four fragments: acetone, formaldehyde, pyruvaldehyde, and malonaldehyde.



- How many double bonds does  $\beta$ -ocimene have?
- Is  $\beta$ -ocimene conjugated or nonconjugated?
- Propose a structure for  $\beta$ -ocimene consistent with the observed data.
- Formulate all of the reactions, showing starting material and products.

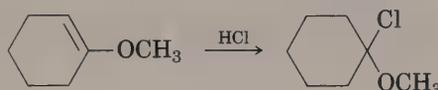
- 14.38 Myrcene,  $C_{10}H_{16}$ , is found in oil of bay leaves and is isomeric with  $\beta$ -ocimene (Problem 14.37). It shows an ultraviolet absorption at 226 nm and can be catalytically hydrogenated to yield 2,6-dimethyloctane. On ozonolysis followed by zinc/acetic acid treatment, myrcene yields formaldehyde, acetone, and 2-oxopentanedial.



2-Oxopentanedial

Propose a structure for myrcene, and formulate all of the reactions, showing starting material and products.

- 14.39 Addition of HCl to 1-methoxycyclohexene yields 1-chloro-1-methoxycyclohexane as the sole product. Why is none of the other regioisomer formed?



- 14.40 Benzene has an ultraviolet absorption at  $\lambda_{\max} = 204$  nm, and *para*-toluidine has  $\lambda_{\max} = 235$  nm. How do you account for this difference?

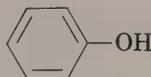


**Benzene**  
( $\lambda_{\max} = 204$  nm)



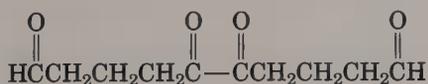
***para*-Toluidine**  
( $\lambda_{\max} = 235$  nm)

- 14.41 When the ultraviolet spectrum of *para*-toluidine (Problem 14.40) is measured in the presence of a small amount of HCl, the  $\lambda_{\max}$  decreases to 207 nm, nearly the same value as for benzene. How do you account for the effect of acid?
- 14.42 Phenol is a weak acid with  $pK_a = 10.0$ . In ethanol solution, phenol has an ultraviolet absorption at  $\lambda_{\max} = 210$  nm. When dilute NaOH is added to this solution, the absorption increases to  $\lambda_{\max} = 235$  nm. How do you account for this shift?

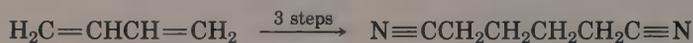


**Phenol**

- 14.43 Hydrocarbon A,  $C_{10}H_{14}$ , has an ultraviolet absorption at  $\lambda_{\max} = 236$  nm and gives hydrocarbon B,  $C_{10}H_{18}$ , on catalytic hydrogenation. Ozonolysis of A followed by zinc/acetic acid treatment yields the following diketo dialdehyde:

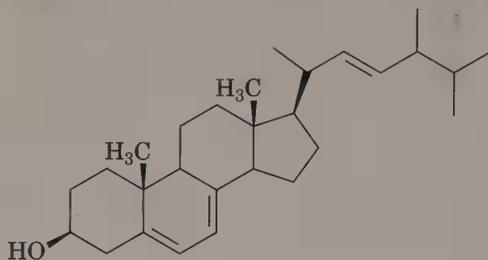


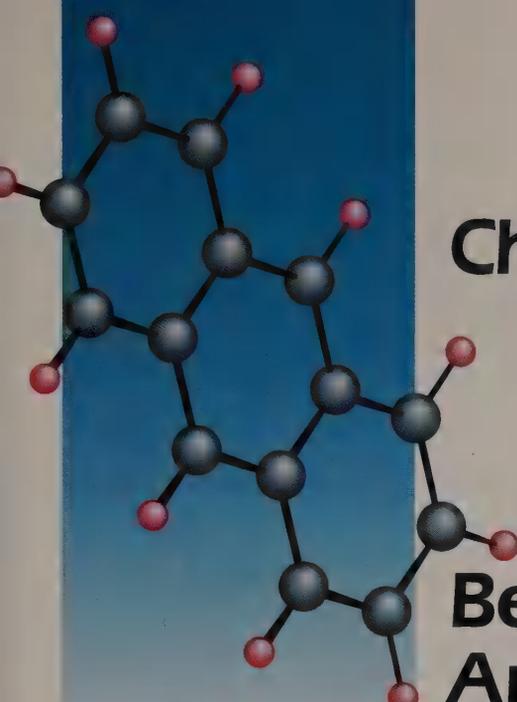
- (a) Propose two possible structures for A.  
 (b) Hydrocarbon A reacts with maleic anhydride to yield a Diels–Alder adduct. Which of your structures for A is correct?  
 (c) Formulate all reactions showing starting material and products.
- 14.44 Adiponitrile, a starting material used in the manufacture of nylon, can be prepared in three steps from 1,3-butadiene. How would you carry out this synthesis?



Adiponitrile

- 14.45 Ergosterol, a precursor of vitamin D, has  $\lambda_{\text{max}} = 282 \text{ nm}$  and molar absorptivity  $\epsilon = 11,900$ . What is the concentration of ergosterol in a solution whose absorbance  $A = 0.065$  with a sample path length  $l = 1.00 \text{ cm}$ ?

Ergosterol ( $\text{C}_{28}\text{H}_{44}\text{O}$ )



## Chapter 15

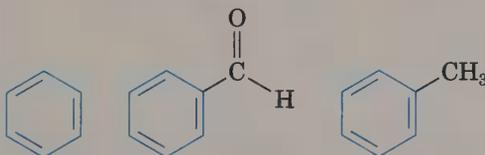
# Benzene and Aromaticity

In the early days of organic chemistry, the word *aromatic* was used to describe fragrant substances such as benzaldehyde (from cherries, peaches, and almonds), toluene (from Tolu balsam), and benzene (from coal distillate). It was soon realized, however, that substances grouped as aromatic behaved in a chemically different manner from most other organic compounds.

Today, we use the term **aromatic** to refer to benzene and its structural relatives. We'll see in this and the next chapter that aromatic substances show chemical behavior quite different from that of the aliphatic substances we've studied to this point. Thus, chemists of the early nineteenth century were correct when they realized that a chemical difference exists between aromatic compounds and others, but the association of aromaticity with fragrance has long been lost.

Many compounds isolated from natural sources are aromatic in part. In addition to benzene, benzaldehyde, and toluene, complex compounds such as the female steroidal hormone, estrone, and the well-known analgesic, morphine, have aromatic rings. Many synthetic drugs used medicinally are also aromatic in part; the tranquilizer diazepam (Valium) is one of many examples.

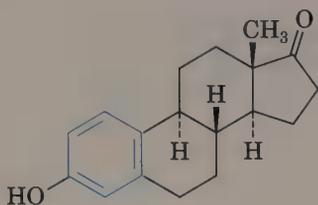
Benzene itself has been found to cause bone-marrow depression and consequent leukopenia (depressed white-blood-cell count) on prolonged exposure. Use of benzene as a laboratory solvent should therefore be avoided.



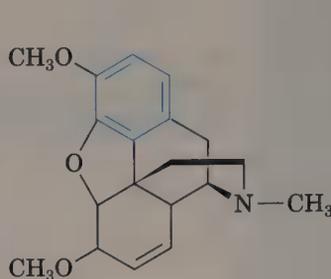
Benzene

Benzaldehyde

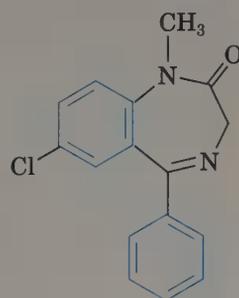
Toluene



Estrone



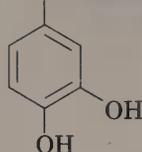
Morphine



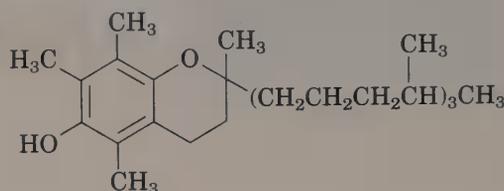
Diazepam (Valium)

**PROBLEM** .....

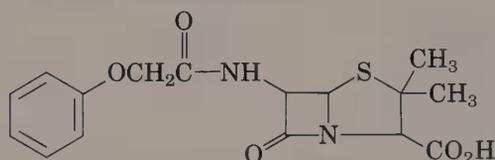
15.1 Circle the aromatic portions of these molecules:



Adrenaline (epinephrine)



Vitamin E



Penicillin V

## 15.1 Sources of Aromatic Hydrocarbons

Simple aromatic hydrocarbons used as starting materials for the preparation of more complex products come from two main sources, coal and petroleum. Coal is an enormously complex substance made up primarily of large arrays of unsaturated benzene-like rings linked together. When heated to  $1000^\circ\text{C}$  in the absence of air, thermal breakdown of coal molecules occurs and a mixture of volatile products, called *coal tar*, distills off. Further fractional distillation of coal tar yields benzene, toluene, xylene (dimethylbenzene), naphthalene, and a host of other aromatic compounds (Figure 15.1).

Petroleum, unlike coal, consists largely of alkanes and contains few aromatic compounds. During petroleum refining (Section 3.6), however, aromatic molecules are formed when alkanes are passed over a catalyst at about  $500^\circ\text{C}$  under high pressure. Heptane ( $\text{C}_7\text{H}_{16}$ ), for example, is converted into toluene ( $\text{C}_7\text{H}_8$ ) by dehydrogenation and cyclization.

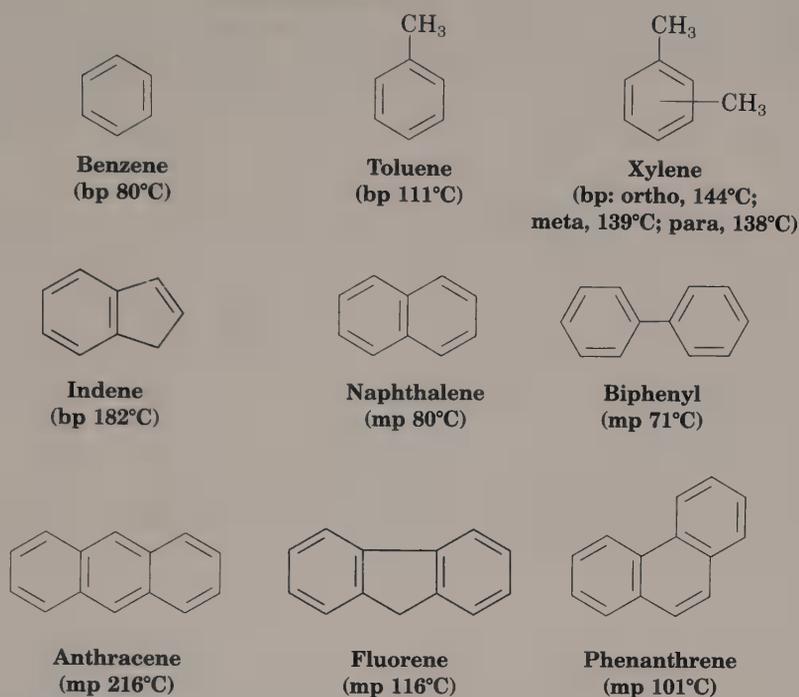


Figure 15.1 Some aromatic hydrocarbons found in coal tar.

## 15.2 Nomenclature of Aromatic Compounds

Aromatic substances, more than any other class of organic compounds, have acquired a large number of nonsystematic names. Although the use of such names is discouraged, IUPAC rules allow for some of the more widely used ones to be retained (Table 15.1). Thus, methylbenzene is known familiarly as *toluene*, hydroxybenzene as *phenol*, aminobenzene as *aniline*, and so on.

Monosubstituted benzene derivatives are systematically named in the same manner as other hydrocarbons, with *-benzene* used as the parent name. Thus, C<sub>6</sub>H<sub>5</sub>Br is bromobenzene, C<sub>6</sub>H<sub>5</sub>NO<sub>2</sub> is nitrobenzene, and C<sub>6</sub>H<sub>5</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> is propylbenzene.

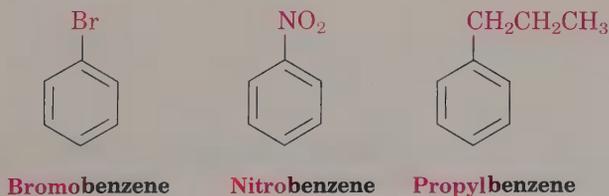
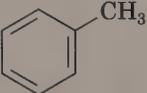
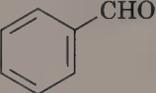
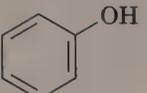
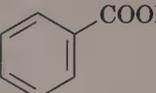
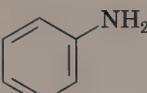
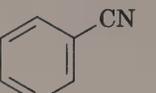
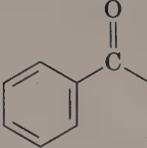
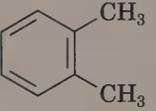
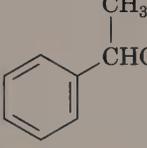
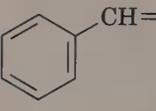
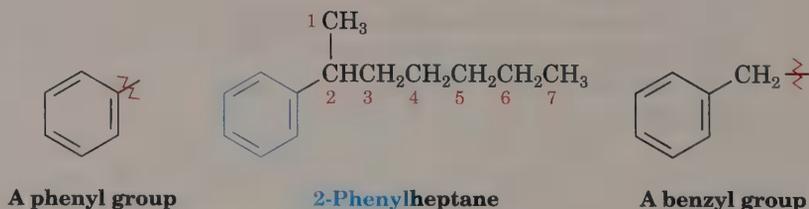


TABLE 15.1 Common Names of Some Aromatic Compounds

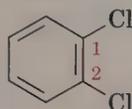
Formula	Name	Formula	Name
	Toluene (bp 110°C)		Benzaldehyde (bp 178°C)
	Phenol (mp 43°C)		Benzoic acid (mp 122°C)
	Aniline (bp 184°C)		Benzonitrile (bp 191°C)
	Acetophenone (mp 21°C)		<i>ortho</i> -Xylene (bp 144°C)
	Cumene (bp 152°C)		Styrene (bp 145°C)

Alkyl-substituted benzenes, sometimes referred to as **arenes**, are named in two different ways depending on the size of the alkyl group. If the alkyl substituent has six or fewer carbons, the arene is named as an alkyl-substituted benzene. If the alkyl substituent is larger than the ring (more than six carbons), the compound is named as a phenyl-substituted alkane. The name **phenyl**, pronounced **fen-nil** and often abbreviated as -Ph or  $\Phi$  (Greek phi), is used for the  $-\text{C}_6\text{H}_5$  unit when the benzene ring is considered as a substituent group. The word is derived from the Greek *pheno* (“I bear light”), commemorating the fact that benzene was discovered by Michael Faraday<sup>1</sup> in 1825 from the oily residue left by the illuminating gas used in London street lamps. As mentioned previously, the  $\text{C}_6\text{H}_5\text{CH}_2-$  alkyl group is called *benzyl*.

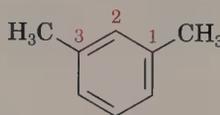


<sup>1</sup>Michael Faraday (1791–1867); b. Newington Butts, Surrey, England; assistant to Sir Humphry Davy (1813); director, laboratory of the Royal Institution (1825); Fullerian Professor of Chemistry, Royal Institution (1833).

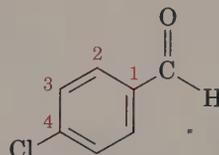
Disubstituted benzenes are named using one of the prefixes *ortho*-, *meta*-, or *para*-. An *ortho*- or *o*-disubstituted benzene has the two substituents in a 1,2 relationship on the ring; a *meta* or *m*-disubstituted benzene has its two substituents in a 1,3 relationship; and a *para*- or *p*-disubstituted benzene has its substituents in a 1,4 relationship.



***ortho*-Dichlorobenzene**  
1,2-Disubstituted

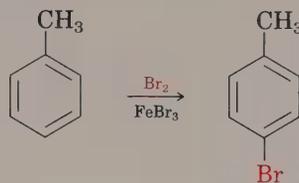
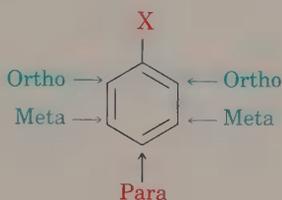


***meta*-Xylene**  
1,3-Disubstituted



***para*-Chlorobenzaldehyde**  
1,4-Disubstituted

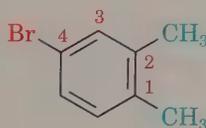
The *ortho*, *meta*, *para* system of nomenclature is also valuable when discussing reactions. For example, we might describe the reaction of bromine with toluene by saying, "Reaction occurs in the *para* position"—in other words, at the position *para* to the methyl group already present on the ring.



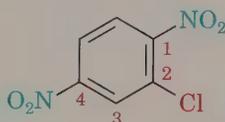
**Toluene**

***p*-Bromotoluene**

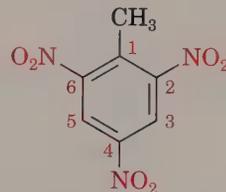
Benzenes with more than two substituents are named by numbering the position of each substituent on the ring. The numbering is carried out such that the lowest possible numbers are used, and the substituents are listed alphabetically when writing the name.



**4-Bromo-1,2-dimethylbenzene**

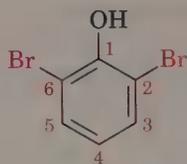


**2-Chloro-1,4-dinitrobenzene**

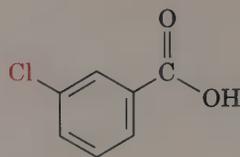


**2,4,6-Trinitrotoluene (TNT)**

Note in the third example shown that *-toluene*, rather than *-benzene*, is used as the base name. Any of the monosubstituted aromatic compounds shown in Table 15.1 can serve as a base name. In such cases the principal substituent ( $-\text{CH}_3$  in toluene, for example) is assumed to be on C1. The following two examples further illustrate this rule, with the base name italicized in each case.

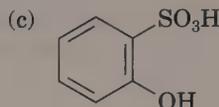
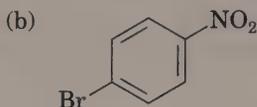
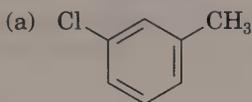


2,6-Dibromophenol

*m*-Chlorobenzoic acid

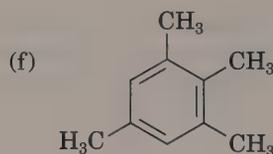
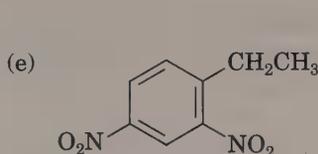
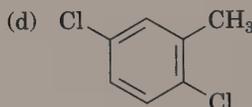
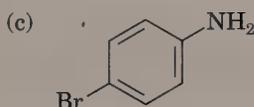
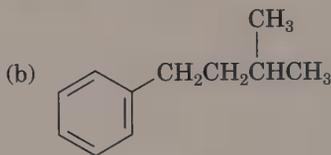
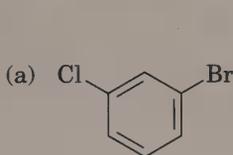
## PROBLEM .....

15.2 Tell whether these compounds are ortho, meta, or para substituted:



## PROBLEM .....

15.3 Provide correct IUPAC names for these compounds:



## PROBLEM .....

15.4 Draw structures corresponding to these IUPAC names:

(a) *p*-Bromochlorobenzene(b) *p*-Bromotoluene(c) *m*-Chloroaniline

(d) 1-Chloro-3,5-dimethylbenzene

## PROBLEM .....

15.5 The following names are incorrect. Draw the structure represented by each, and provide correct IUPAC names.

(a) 2-Bromo-3-chlorobenzene

(b) 4,6-Dinitrotoluene

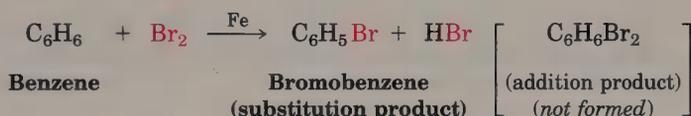
(c) 4-Bromo-1-methylbenzene

(d) 2-Chloro-*p*-xylene

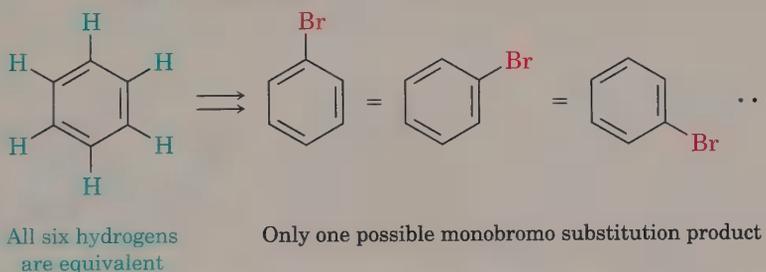
## .....

## 15.3 Structure of Benzene: The Kekulé Proposal

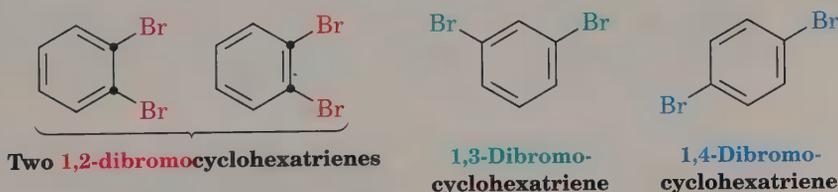
By the mid-1800s, benzene was known to have the molecular formula  $C_6H_6$ , and its chemistry was being actively explored. The results, though, were puzzling. Although benzene is clearly “unsaturated”—the formula  $C_6H_6$  requires a combination of four double bonds/rings—it nevertheless fails to undergo reactions characteristic of alkenes. For example, it reacts slowly with bromine in the presence of iron to give the *substitution* product  $C_6H_5Br$ , rather than the possible *addition* product  $C_6H_6Br_2$ . Furthermore, only one monobromo substitution product was known; no isomers of  $C_6H_5Br$  had been prepared.



On further reaction of bromobenzene with bromine, disubstitution products are obtained, and three isomeric  $C_6H_4Br_2$  compounds had been prepared. On the basis of these and similar results, August Kekulé proposed in 1865 that benzene consists of a *ring* of carbon atoms and can be formulated as 1,3,5-cyclohexatriene. Kekulé reasoned that this structure would readily account for the isolation of only a single monobromo substitution product, since all six carbon atoms and all six hydrogens in 1,3,5-cyclohexatriene are equivalent.

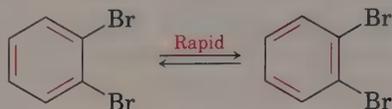


The observation that only three isomeric dibromo substitution products were known was more difficult to explain, since four structures can be written:



Although there is only one possible 1,3 derivative and one possible 1,4 derivative, there appear to be two possible 1,2-dibromo substitution products, depending on the positions of the double bonds in the ring. One isomer

has a single bond between the bromine-bearing carbons, and the other isomer has a double bond. Kekulé accounted for the formation of only three isomers by proposing that the double bonds in benzene rapidly “oscillate” between two positions. Thus, the two “1,2-dibromocyclohexatrienes” can’t be separated, according to Kekulé, because they interconvert too rapidly.



Kekulé’s proposed structure for benzene was widely criticized at the time. Although it satisfactorily accounts for the correct number of mono- and disubstituted benzene isomers, it fails to answer two critical questions: Why is benzene unreactive compared to other alkenes, and why does benzene give a substitution product rather than an addition product on reaction with bromine?

PROBLEM .....

- 15.6 How many tribromo benzene derivatives are possible according to Kekulé’s theory? Draw and name them.

PROBLEM .....

- 15.7 The following structures with formula  $C_6H_6$  were at one time suggested for benzene. If we assume that bromine can be substituted for hydrogen in these structures, how many monobromo derivatives are possible for each? How many dibromo derivatives?



Ladenburg benzene

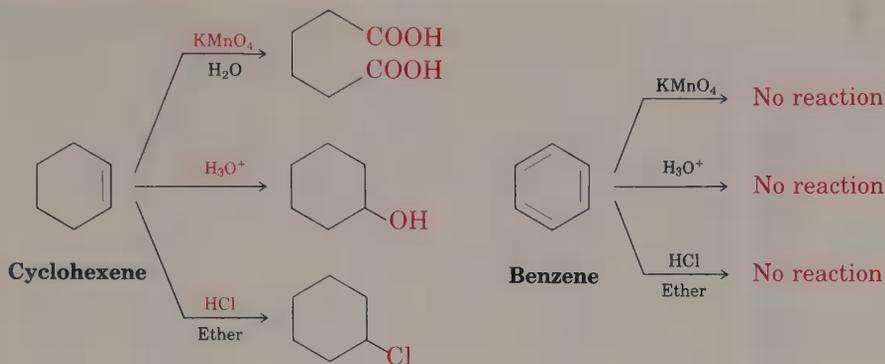


Dewar benzene

## 15.4 Stability of Benzene

The unusual stability of benzene was a great puzzle to early chemists. Although its formula,  $C_6H_6$ , indicates that several double bonds must be present, benzene shows none of the behavior characteristic of alkenes. For example, alkenes react readily with potassium permanganate to give cleavage products, they react with aqueous acid to give alcohols, and they react with HCl to give saturated alkyl chlorides. Benzene does none of these things (Figure 15.2). *Benzene does not undergo electrophilic addition reactions.*

We can get a quantitative idea of benzene’s unusual stability by looking at data on heats of hydrogenation (Table 15.2). Cyclohexene, an isolated alkene, has  $\Delta H^\circ_{\text{hydrog}} = 28.6 \text{ kcal/mol}$  (120 kJ/mol), and 1,3-cyclohexadiene, a conjugated diene, has  $\Delta H^\circ_{\text{hydrog}} = 55.4 \text{ kcal/mol}$  (232 kJ/mol). As expected, the value for 1,3-cyclohexadiene is a bit less than twice the cyclohexene value since conjugated dienes are unusually stable (Section 14.2).



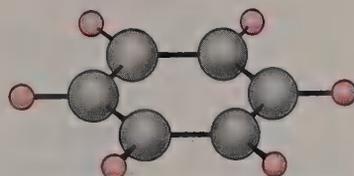
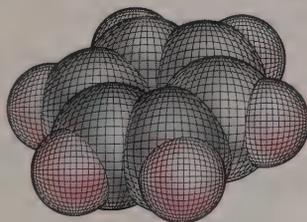
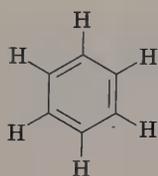
**Figure 15.2** A comparison of the reactivity of cyclohexene and benzene. Benzene does not undergo electrophilic addition reactions.

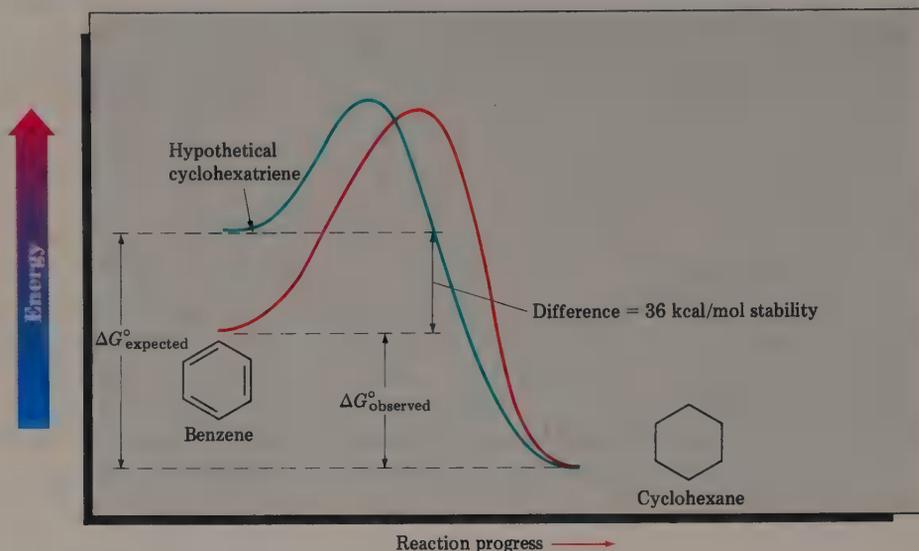
**TABLE 15.2** Heats of Hydrogenation of Cyclic Alkenes

Reactant	Product	$\Delta H^\circ_{\text{hydrog}}$	
		(kcal/mol)	(kJ/mol)
Cyclohexene	Cyclohexane	28.6	120
1,3-Cyclohexadiene	Cyclohexane	55.4	232
Benzene	Cyclohexane	49.8	208

Carrying the analogy one step further, we might predict that  $\Delta H^\circ_{\text{hydrog}}$  for “cyclohexatriene” (benzene) should be about 86 kcal/mol, or three times the cyclohexene value. *The actual value is 49.8 kcal/mol (208 kJ/mol), some 36 kcal/mol less than expected.* Since 36 kcal/mol less heat than expected is released during hydrogenation of benzene, benzene must have 36 kcal/mol less energy than expected. In other words, benzene has 36 kcal/mol “extra” stability (Figure 15.3).

Further evidence for the unusual nature of benzene comes from spectroscopic studies showing that all carbon–carbon bonds in benzene have the same length, intermediate between a normal single and a normal double bond. Although most C–C single bonds are 1.54 Å long and most C–C double bonds are 1.34 Å long, the carbon–carbon bonds in benzene are 1.39 Å long.

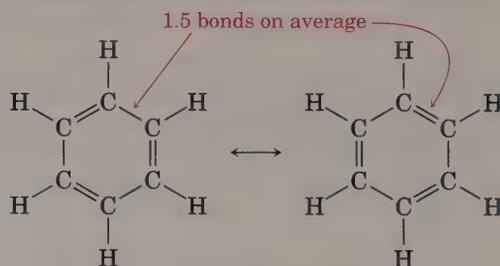




**Figure 15.3** Reaction energy diagram for the hydrogenation of benzene in comparison with a hypothetical cyclohexatriene.

## 15.5 Representations of Benzene: The Resonance Approach

How can we account for benzene's properties, and how can we best represent its structure? Resonance theory answers this question by saying that a *single* Kekulé structure is not satisfactory and that benzene can best be described as a resonance hybrid of two equivalent Kekulé structures in which each carbon–carbon bond averages out to be 1.5 bonds—midway between single and double.



Recall from Section 10.7 that resonance theory has four postulates:

1. Resonance forms are imaginary, not real. The real structure of benzene is a single, unchanging hybrid structure that combines the characteristics of both resonance forms.

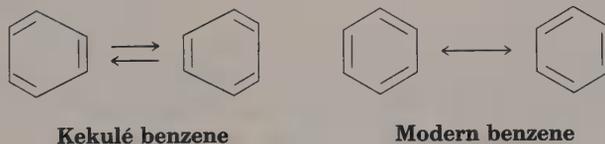
2. Resonance structures differ only in the positions of their electrons. Neither the position nor the hybridization of atoms changes from one resonance structure to another. In benzene, the six carbon nuclei form a regular hexagon while the pi electrons are shared equally between neighboring nuclei. Each carbon–carbon bond averages 1.5 electrons, and all bonds are equivalent.
3. Different resonance forms don't have to be equivalent. The more nearly equivalent the forms are, however, the more stable the molecule. Thus benzene, with two equivalent resonance forms, is highly stable.
4. The more resonance structures there are, the more stable the molecule.

The true benzene structure can't be represented accurately by either single Kekulé structure and does not oscillate back and forth between the two. The true structure is somewhere in between the two extremes but is impossible to draw with our usual conventions. We might try to represent benzene by drawing it with either a full or a dotted circle to indicate the equivalence of all carbon–carbon bonds, but these representations have to be used very carefully since they don't indicate the number of pi electrons in the ring—how many electrons does a circle represent? In this book, benzene and other arenes will be represented by a single Kekulé structure. We'll be able to keep count of all pi electrons this way, but we must keep aware of the limitations of the drawings.

Some alternative representations of benzene. Such structures must be used carefully, since they don't indicate the number of pi electrons.



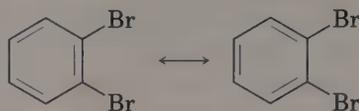
There's a subtle yet important difference between Kekulé's representation of benzene and the resonance representation. Kekulé considered benzene as rapidly oscillating back and forth between two cyclohexatriene structures, whereas resonance theory considers benzene to be a single "resonance hybrid" structure:



At any given instant, Kekulé's oscillating structures don't have all their carbon–carbon bonds the same length: Three bonds are short and three are long. This difference in bond length implies that the carbon atoms must change position in oscillating from one structure to another, and thus the two structures are not the same as resonance forms.

The same resonance argument that explains benzene's stability also explains why there's only one *ortho*-dibromobenzene, rather than two. The two Kekulé structures of *o*-dibromobenzene are simply different resonance

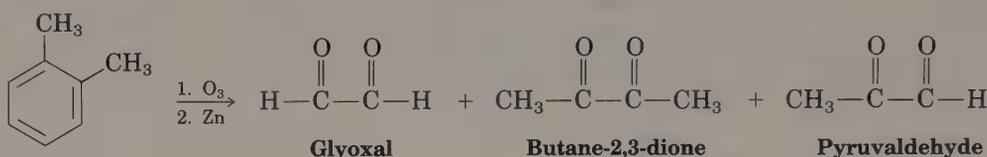
forms of a single compound whose true structure is intermediate between the two.



Two equivalent resonance forms of *o*-dibromobenzene

PROBLEM

- 15.8 In 1932, A. A. Levine and A. G. Cole studied the ozonolysis of *o*-xylene. They isolated three products: glyoxal (OHC-CHO), butane-2,3-dione (CH<sub>3</sub>COCOCH<sub>3</sub>), and pyruvaldehyde (CH<sub>3</sub>COCHO):



In what ratio would you expect these three products to be formed if *o*-xylene is a resonance hybrid of two Kekulé structures? The actual ratio found was 3 parts glyoxal, 2 parts pyruvaldehyde, and 1 part butane-2,3-dione. What conclusions can you draw about the structure of *o*-xylene?

## 15.6 Molecular Orbital Description of Benzene

Having just seen a resonance description of benzene, let's now see a molecular orbital description. MO theory describes benzene in a way that is, in many respects, superior to the simple resonance approach.

The orbital view of benzene in Figure 15.4 emphasizes the cyclic conjugation of the benzene molecule and the equivalence of the six carbon-carbon bonds. Benzene is a flat, symmetrical molecule with the shape of a regular hexagon. All C-C-C bond angles are 120°, all six carbon atoms are *sp*<sup>2</sup> hybridized, and each carbon has a *p* orbital perpendicular to the plane of the six-membered ring.

Since all six carbon atoms and all six *p* orbitals in benzene are equivalent, it's impossible to define three localized alkene pi bonds in which a given *p* orbital overlaps only one neighboring *p* orbital. Rather, each *p* orbital overlaps equally well with *both* neighboring *p* orbitals, leading to a picture of benzene in which the six pi electrons are completely delocalized around the ring. Benzene therefore has two doughnut-shaped clouds of electrons, one above and one below the ring.

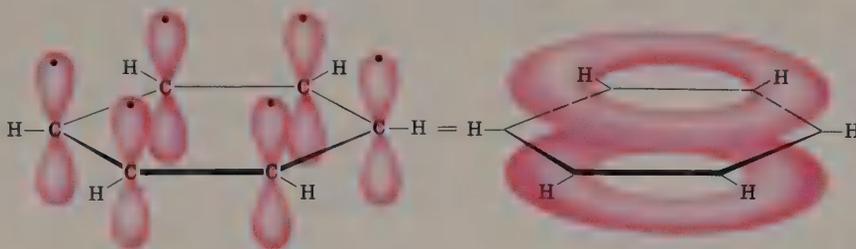


Figure 15.4 An orbital picture of benzene.

We can construct molecular orbitals for benzene just as we did for 1,3-butadiene in the preceding chapter. If we allow six  $p$  atomic orbitals to combine in a cyclic manner, six benzene molecular orbitals result, as shown in Figure 15.5. The three low-energy molecular orbitals, denoted  $\psi_1$ ,  $\psi_2$ , and  $\psi_3$ , are lower in energy than the  $p$  atomic orbitals and are therefore bonding combinations. The three high-energy orbitals are antibonding. Note that two of the bonding orbitals,  $\psi_2$  and  $\psi_3$ , have the same energy, as do the antibonding orbitals  $\psi_4^*$  and  $\psi_5^*$ . Such orbitals are said to be **degenerate**. The six  $p$  electrons of benzene occupy the three bonding molecular orbitals and are therefore delocalized over the entire conjugated system, leading to the observed 36 kcal/mol stabilization energy of benzene.

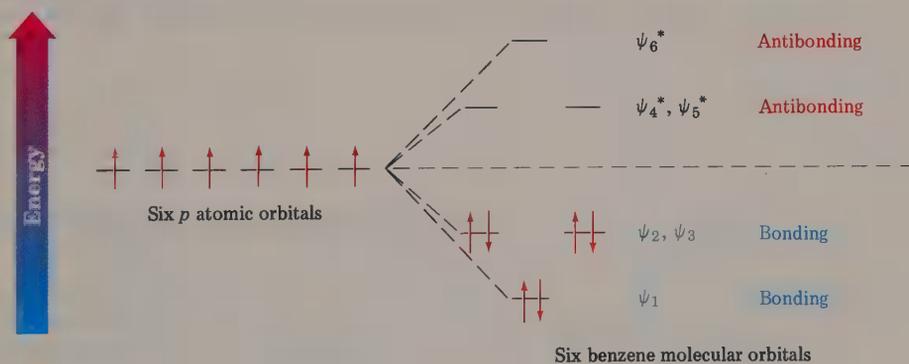
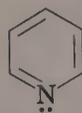


Figure 15.5 The six molecular orbitals of benzene.

PROBLEM .....

15.9 Pyridine is a flat, symmetrical molecule with bond angles of  $120^\circ$ . It undergoes electrophilic substitution rather than addition and generally behaves as an aromatic

substance. Draw an orbital picture of pyridine to explain its properties. Check your answer by looking ahead to Section 15.9.



Pyridine

## 15.7 Aromaticity and the Hückel $4n + 2$ Rule

Let's review what we've learned thus far about benzene and, by extension, about other benzene-like (**benzenoid**) aromatic molecules:

1. Benzene is a monocyclic conjugated molecule with formula  $C_6H_6$ .
2. Benzene is unusually stable, having a heat of hydrogenation 36 kcal/mol less than we might expect for a cyclic conjugated triene.
3. Benzene is a symmetrical, planar, hexagonal molecule. All C–C bond angles are  $120^\circ$ , and all C–C bonds are 1.39 Å in length.
4. Benzene undergoes substitution reactions that retain the cyclic conjugation, rather than electrophilic addition reactions that would destroy the conjugation.
5. Benzene can be described in terms of resonance theory as a hybrid whose structure is intermediate between two Kekulé structures:



Taken together, these facts provide a description of benzene and of other aromatic molecules. But they aren't sufficient. Something else is needed to complete a description of aromaticity. According to theoretical calculations carried out by the German physicist Erich Hückel<sup>2</sup> in 1931, a molecule can be aromatic only if it has a planar, monocyclic system of conjugation with a  $p$  orbital on each atom and *only if the  $p$  orbital system contains  $4n + 2$  pi electrons*, where  $n$  is an integer ( $n = 0, 1, 2, 3, \dots$ ). In other words, only molecules with 2, 6, 10, 14, 18, . . . pi electrons can be aromatic. Molecules with  $4n$  pi electrons (4, 8, 12, 16, . . .) *can't* be aromatic, even though they may be cyclic and conjugated.

<sup>2</sup>Erich Hückel (1896–1980); b. Charlottenburg, Germany; Ph.D. Göttingen (Debye); professor of physics at Stuttgart and Marburg.

Let's look at some examples to see how the **Hückel  $4n + 2$  rule** works.

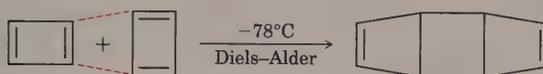
1. *Cyclobutadiene* has four pi electrons and is not aromatic:



Two double bonds;  
four pi electrons

**Cyclobutadiene**

Cyclobutadiene is a highly reactive molecule that shows none of the properties associated with aromaticity. A long history of attempts to synthesize the compound culminated in 1965 when Rowland Pettit<sup>3</sup> of the University of Texas prepared cyclobutadiene at low temperature but was unable to isolate it. Even at  $-78^\circ\text{C}$ , however, cyclobutadiene dimerizes by Diels–Alder cycloaddition. One molecule behaves as a diene, and the other acts as dienophile:



2. *Benzene* has six pi electrons ( $4n + 2$ , where  $n = 1$ ) and is, of course, a fully aromatic molecule:



Three double bonds;  
six pi electrons

**Benzene**

3. *Cyclooctatetraene* has eight pi electrons and is not aromatic:



Four double bonds;  
eight pi electrons

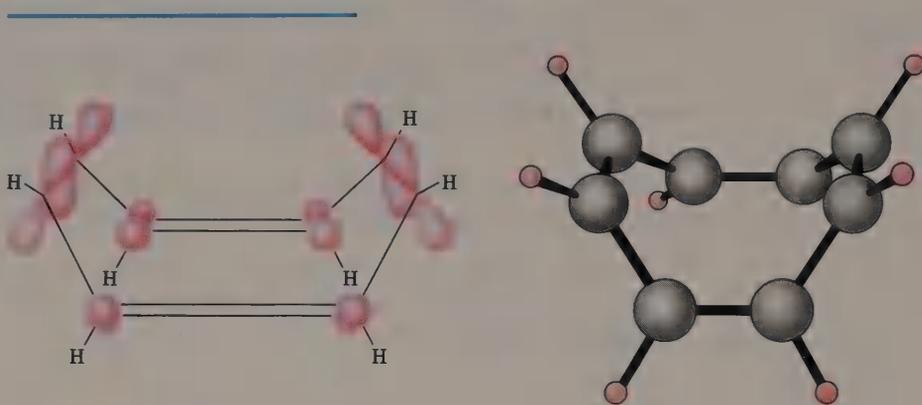
**Cyclooctatetraene**

In the early 1900s, chemists believed that the only requirement for aromaticity was the presence of a cyclic conjugated system. It was therefore expected that cyclooctatetraene, as a close analog of benzene, would also prove to be unusually stable. The facts proved otherwise. When cyclooctatetraene was first prepared in 1911 by the German chemist Richard Willstätter,<sup>4</sup> it was found to resemble open-chain polyenes in its reactivity.

<sup>3</sup>Rowland Pettit (1927–1981); b. Port Lincoln, Australia; Ph.D. University of Adelaide (1952) and University of London (Dewar, 1956); University of Texas (1957–1981).

<sup>4</sup>Richard Willstätter (1872–1942); b. Karlsruhe, Germany; Technische Hochschule, Munich (Einhorn) (1895); professor, Zürich, Dahlem, Munich; Nobel Prize (1915).

Cyclooctatetraene reacts readily with bromine, with potassium permanganate, and with hydrogen chloride, just as other alkenes do. We now know, in fact, that cyclooctatetraene is not even conjugated. It is tub shaped, rather than planar, and has no cyclic conjugation because neighboring orbitals don't have the proper geometry for overlap (Figure 15.6). The pi electrons are localized in four distinct carbon-carbon double bonds rather than delocalized as in benzene. X-ray studies show that the carbon-carbon single bonds are 1.47 Å long, whereas the double bonds are 1.34 Å long. In addition, the  $^1\text{H}$  NMR spectrum shows a single sharp resonance line at 5.7  $\delta$ , characteristic of an alkene rather than an aromatic molecule (Section 15.12).



**Figure 15.6** An orbital view of cyclooctatetraene, a tub-shaped molecule that has no cyclic conjugation because its  $p$  orbitals are not aligned properly for overlap.

**PROBLEM** .....

- 15.10** In order to be aromatic, a molecule must be flat, so that  $p$ -orbital overlap can occur, and must have  $4n + 2$  pi electrons. Cyclodecapentaene fulfills one of these criteria but not the other and is therefore nonaromatic. Explain. (Molecular models may be useful.)

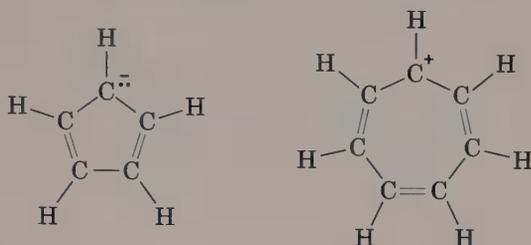


Cyclodecapentaene (not aromatic)

## 15.8 Aromatic Ions

Look back at the Hückel criteria for aromaticity in the preceding section. In order to be aromatic, a molecule must have a  $p$  orbital on each atom and must have  $4n + 2$  pi electrons. There's nothing in this definition that says

the number of  $p$  orbitals and the number of pi electrons have to be the same. In fact, they can be different. The  $4n + 2$  rule is broadly applicable to many kinds of molecules, not just to neutral hydrocarbons. For example, both the cyclopentadienyl *anion* and the cycloheptatrienyl *cation* are aromatic.



Cyclopentadienyl anion      Cycloheptatrienyl cation

Six pi electrons; aromatic ions

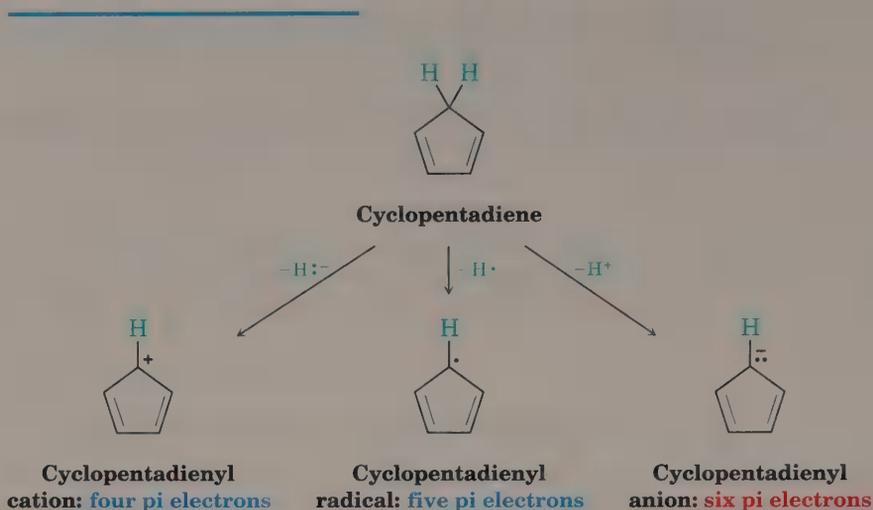
Let's look first at the cyclopentadienyl anion. Cyclopentadiene itself is not aromatic because it's not a fully conjugated molecule: The  $-\text{CH}_2-$  carbon in the ring is  $sp^3$  hybridized, thus preventing complete cyclic  $p$ -orbital conjugation. Imagine, though, that we remove one hydrogen from the saturated  $\text{CH}_2$  group and let that carbon become  $sp^2$  hybridized. The resultant species would have five  $p$  orbitals, one on each of the five carbons, and would be fully conjugated.

There are three ways, shown in Figure 15.7, we might imagine removing the hydrogen:

1. We could remove the hydrogen atom and *both* electrons from the C–H bond. Since the hydrogen removed must carry a negative charge, the cyclopentadienyl group that remains is positively charged.
2. We could remove the hydrogen and *one* electron from the C–H bond, leaving a cyclopentadienyl radical.
3. We could remove the hydrogen ion ( $\text{H}^+$ ) with *no* electrons, leaving a cyclopentadienyl anion.

Resonance theory predicts that all three species should be stabilized, since it's possible to draw five equivalent resonance structures for each. Hückel theory, however, predicts that *only* the six-pi-electron anion should be aromatic. The four-pi-electron cyclopentadienyl carbocation and the five-pi-electron cyclopentadienyl radical are predicted not to be aromatic, even though they have cyclic conjugation.

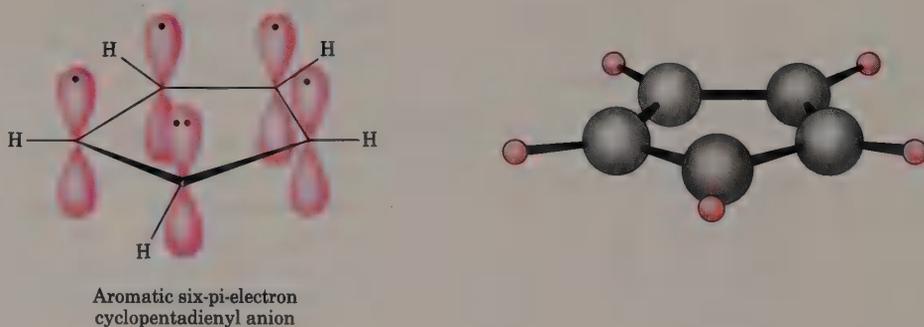
In practice, both the cyclopentadienyl cation and the radical are highly reactive and difficult to prepare; neither species shows any sign of the unusual stability expected of an aromatic system. The six-pi-electron cyclopentadienyl anion, by contrast, is an easily prepared, remarkably stable carbanion (carbon anion). In fact, cyclopentadiene is one of the most acidic hydrocarbons known. Although most hydrocarbons have a  $\text{p}K_a > 45$ , cyclopentadiene has a  $\text{p}K_a$  of 16, a value comparable to that of water! Cyclopentadiene is acidic because the anion formed by ionization is so stable. It doesn't matter that the cyclopentadienyl anion has only five  $p$  orbitals; all that matters is that there are six pi electrons, a Hückel number (Figure 15.8).



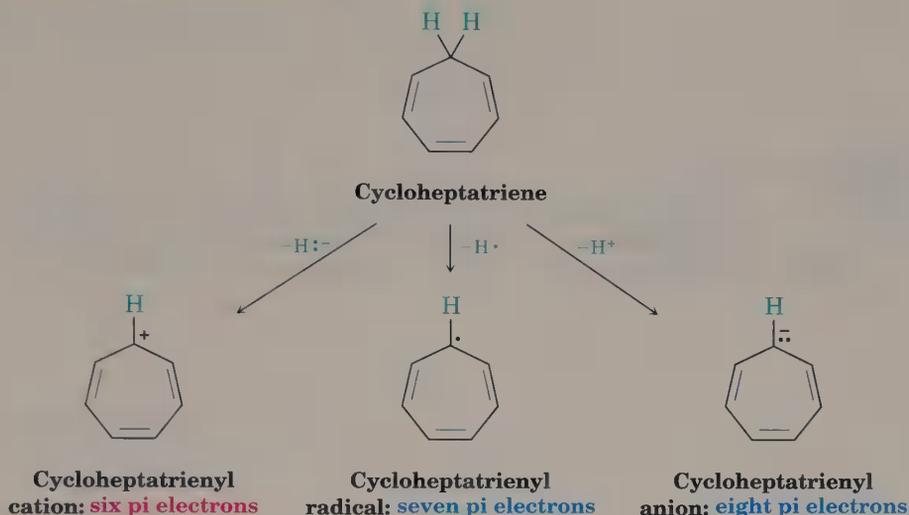
**Figure 15.7** Generating the cyclopentadienyl cation, radical, and anion by removing a hydrogen from cyclopentadiene.

Similar arguments can be used to predict the stabilities of the cycloheptatrienyl cation, radical, and anion. Removal of a hydrogen from cycloheptatriene can generate the six-pi-electron cation, the seven-pi-electron radical, or the eight-pi-electron anion (Figure 15.9). Resonance theory once again predicts a high level of stability for all three species, since seven equivalent resonance forms can be drawn for each. Hückel molecular orbital theory, however, predicts that only the six-pi-electron cycloheptatrienyl cation should have aromatic stability.

Both the seven-pi-electron cycloheptatrienyl radical and the eight-pi-electron anion are highly reactive and difficult to prepare. The six-pi-

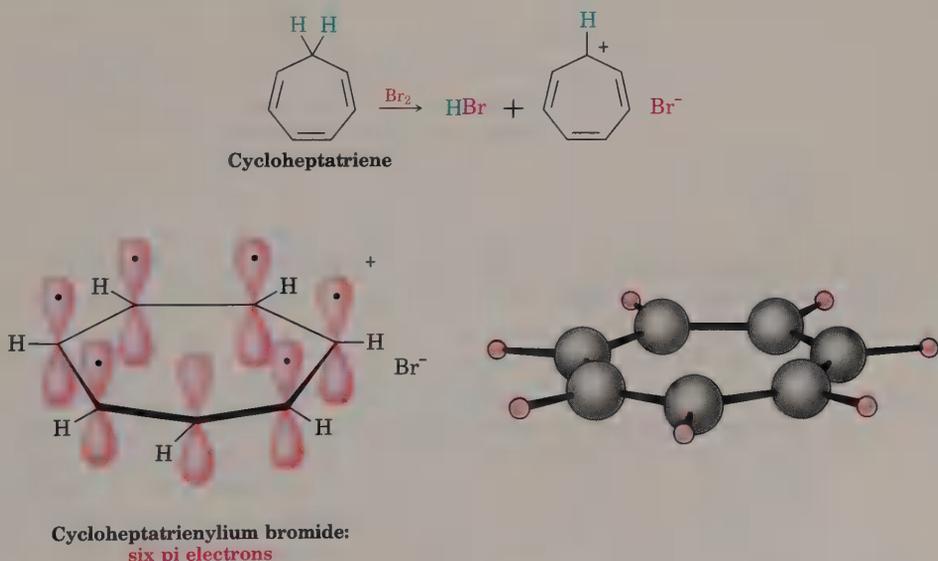


**Figure 15.8** An orbital view of the aromatic cyclopentadienyl anion, showing the cyclic conjugation and six pi electrons in five  $p$  orbitals.



**Figure 15.9** Generation of the cycloheptatrienyl cation, radical, and anion. Only the six-pi-electron cation is aromatic.

electron cation, however, is extraordinarily stable. In fact, the cycloheptatrienyl cation was first prepared in 1891, although its structure was not recognized at the time. Later investigation showed that the salt-like material prepared by action of bromine on cycloheptatriene is cycloheptatrienylium bromide (Figure 15.10).



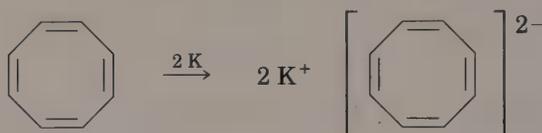
**Figure 15.10** Preparation of the aromatic cycloheptatrienyl cation by reaction of cycloheptatriene with bromine.

## PROBLEM.....

- 15.11 Draw all possible resonance structures of the cyclopentadienyl anion. Are all carbon-carbon bonds equivalent? How many absorption lines would you expect to see in the  $^1\text{H}$  NMR and  $^{13}\text{C}$  NMR spectra of the anion?

## PROBLEM.....

- 15.12 Cyclooctatetraene readily accepts two electrons from potassium metal to form the cyclooctatetraene dianion,  $\text{C}_8\text{H}_8^{2-}$ . Why do you suppose this reaction occurs? What geometry would you expect for the cyclooctatetraene dianion?



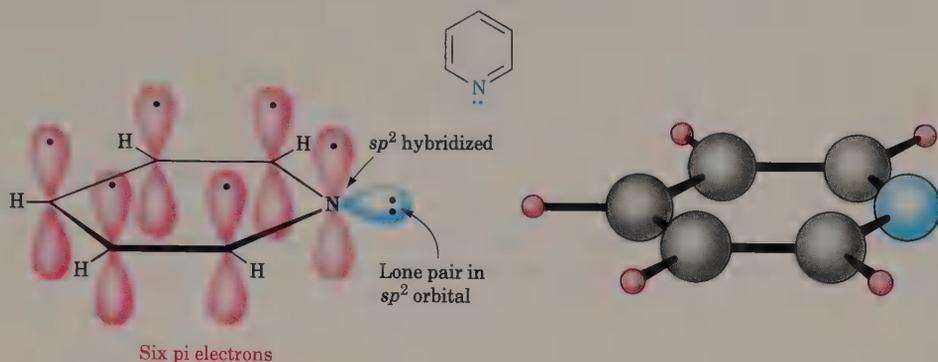
## 15.9 Pyridine and Pyrrole: Two Aromatic Heterocycles

Look back once again at the definition of aromaticity: . . . a molecule that has a cyclic system of conjugation with a  $p$  orbital on each *atom* and that contains  $4n + 2$  pi electrons. Nothing in this definition says that the atoms in the ring have to be *carbon*. In fact, *heterocyclic* compounds can also be aromatic. A **heterocycle** is a cyclic compound that has one or more atoms other than carbon in its ring. The heteroatom is often nitrogen (or oxygen, but sulfur, phosphorus, and other elements are also found. Pyridine, for example, is a six-membered heterocycle with a nitrogen atom in its ring.

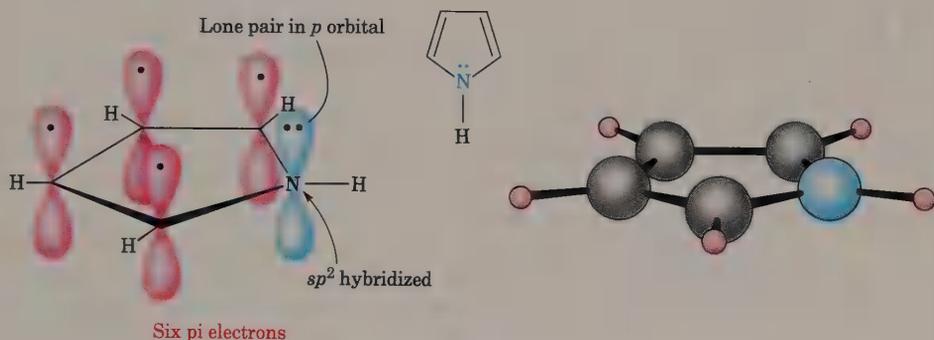
Pyridine is much like benzene in its pi electron structure. Each of the five  $sp^2$ -hybridized carbons has a  $p$  orbital perpendicular to the plane of the ring, and each contains one pi electron. The nitrogen atom also is  $sp^2$  hybridized and has one electron in a  $p$  orbital, bringing the total to six pi electrons. The nitrogen lone-pair electrons are in an  $sp^2$  orbital in the plane of the ring perpendicular to the pi system and are not involved with the aromatic pi system since they don't have the correct geometry for overlap (Figure 15.11).

The five-membered heterocycle, pyrrole (two *r*'s, one *l*), is another example of an aromatic six-pi-electron substance, with a pi electron system similar to that of the cyclopentadienyl anion. Each of the four  $sp^2$ -hybridized carbons has a  $p$  orbital perpendicular to the ring, and each contributes one pi electron. The nitrogen atom is  $sp^2$  hybridized, with its lone pair of electrons also occupying a  $p$  orbital. Thus, there are a total of six pi electrons, making pyrrole an aromatic molecule. An orbital picture of pyrrole is shown in Figure 15.12.

Note that the nitrogen atoms play different roles in pyridine and pyrrole even though both compounds are aromatic. The nitrogen atom in pyridine



**Figure 15.11** Pyridine, an aromatic heterocycle, has a pi electron arrangement much like that of benzene.



**Figure 15.12** Pyrrole, a five-membered aromatic heterocycle, has a pi electron arrangement much like that of the cyclopentadienyl anion.

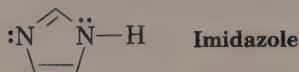
is part of a double bond and therefore contributes only *one* pi electron to the aromatic sextet, just as a carbon atom in benzene does. The nitrogen atom in pyrrole, however, is not part of a double bond. Like one of the carbons in the cyclopentadienyl anion, the pyrrole nitrogen atom contributes *two* pi electrons (the lone pair) to the aromatic sextet.

Pyridine, pyrrole, and other aromatic heterocycles play an important role in many biochemical processes. Their chemistry will be discussed in more detail in Chapter 29.

**PROBLEM** .....

- 15.13** The aromatic five-membered heterocycle imidazole is of great importance in many biological processes. One of its nitrogen atoms is "pyridine-like" in that it contributes one pi electron to the aromatic system, and the other nitrogen is "pyrrole-like" in

that it contributes two pi electrons. Draw an orbital picture of imidazole to account for its aromaticity. How many pi electrons does each nitrogen contribute?



- 15.14 Assume that the oxygen atom in furan is  $sp^2$  hybridized, and draw an orbital picture to show how this heterocycle can be aromatic.



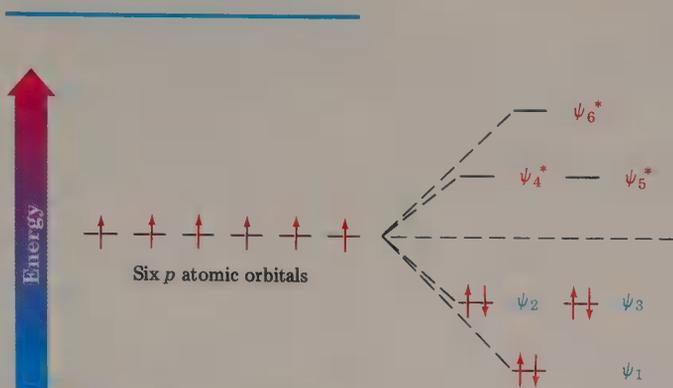
## 15.10 Why $4n + 2$ ?

What's so special about  $4n + 2$  pi electrons? Why is it that having 2, 6, 10, 14, and so on, pi electrons leads to aromatic stability, but having other numbers of electrons does not? The answer to this question has to do with the relative energy levels of the pi molecular orbitals.

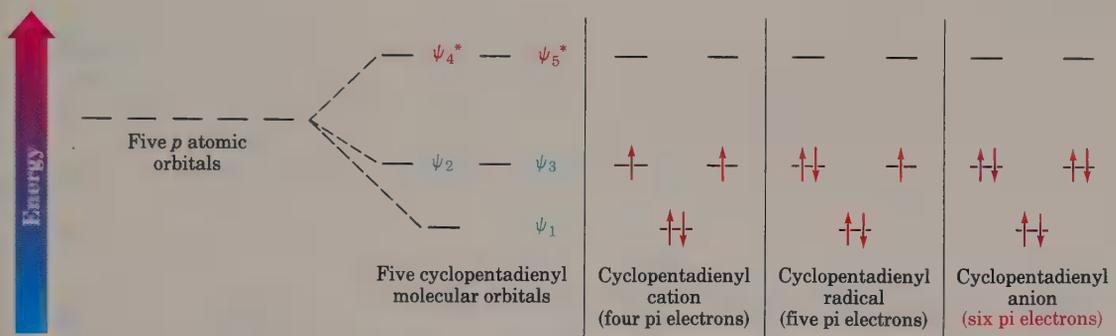
When the energy levels of molecular orbitals for cyclic conjugated molecules are calculated, it turns out that there is always a *single* lowest-lying molecular orbital above which the molecular orbitals come in degenerate pairs. Thus, when electrons fill the various molecular orbitals, it takes two electrons (one pair) to fill the lowest-lying orbital and four electrons (two pairs) to fill each of  $n$  succeeding energy levels. The total is  $4n + 2$ .

This orbital filling is illustrated in Figure 15.13 for benzene, in which six atomic  $p$  orbitals combine to give six molecular orbitals with the energy levels shown. Note that the lowest-energy molecular orbital,  $\psi_1$ , occurs singly and contains two electrons. The next two lowest-energy orbitals,  $\psi_2$  and  $\psi_3$ , are degenerate, and it therefore takes four electrons to fill them. The result is a highly stabilized six-pi-electron aromatic molecule with filled bonding molecular orbitals.

A similar line of reasoning carried out for the cyclopentadienyl cation, radical, and anion leads to the conclusions illustrated in Figure 15.14. Once again, the five atomic  $p$  orbitals combine to give five pi molecular orbitals in which there is a single lowest-energy orbital and higher-energy degenerate pairs of orbitals. In the four-pi-electron cation, there are two electrons in  $\psi_1$  but only one electron each in  $\psi_2$  and  $\psi_3$ . Thus, the cation has two orbitals that are only partially filled and is therefore reactive. In the five-pi-electron radical,  $\psi_1$  and  $\psi_2$  are filled, but  $\psi_3$  is still only half full. The radical is therefore reactive. Only in the six-pi-electron cyclopentadienyl anion are all of the bonding orbitals filled. Similar analyses can be carried out for all other aromatic species.



**Figure 15.13** Energy levels of the six benzene pi molecular orbitals.



**Figure 15.14** Energy levels of the five cyclopentadienyl molecular orbitals. Only the six-pi-electron cyclopentadienyl anion has a filled-shell configuration leading to aromatic stability.

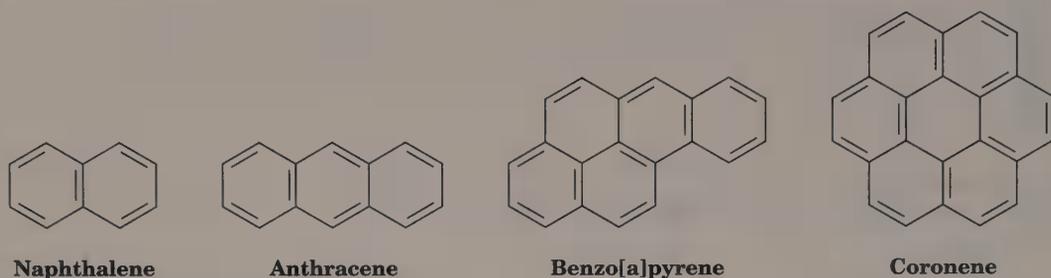
**PROBLEM** .....

- 15.15 Show the relative positions of the seven pi molecular orbitals of the cycloheptatrienyl system. Indicate which of the seven orbitals are filled in the cation, radical, and anion, and account for the aromaticity of the cycloheptatrienyl cation.
- .....

## 15.11 Naphthalene: A Polycyclic Aromatic Compound

Although the Hückel rule is strictly applicable only to *monocyclic* aromatic compounds, the general concept of aromaticity can be extended beyond simple monocyclic compounds like benzene to include **polycyclic aromatic**

**compounds.** Naphthalene, with two benzene-like rings fused together, is the simplest polycyclic aromatic molecule, but more complex substances such as anthracene, 1,2-benzopyrene, and coronene are also well known (Figure 15.15). Benzo[*a*]pyrene is particularly interesting because it's one of the cancer-causing substances that has been isolated from tobacco smoke.



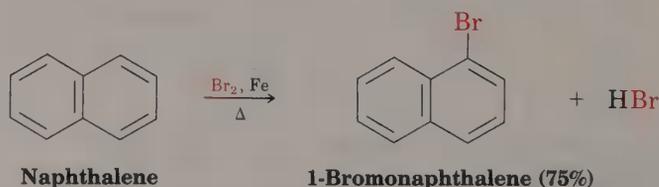
**Figure 15.15** Some polycyclic aromatic hydrocarbons.

All polycyclic aromatic hydrocarbons can be represented by a number of different Kekulé structures. Naphthalene, for instance, has three resonance forms:



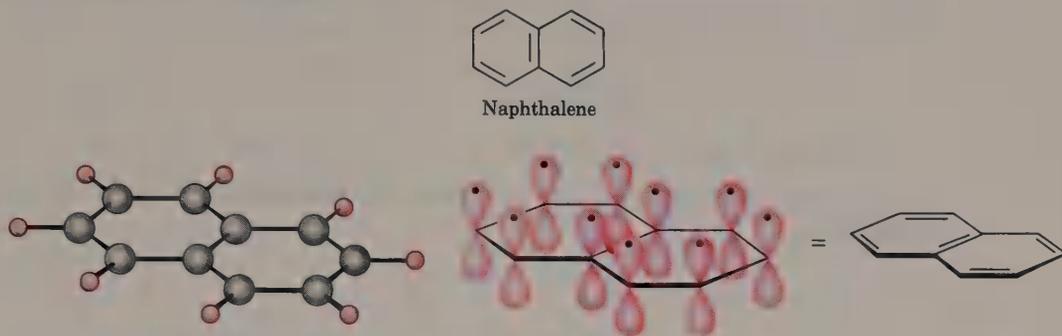
As was true for benzene with its two equivalent resonance forms, no single Kekulé structure is a true representation of naphthalene; the true structure of naphthalene lies somewhere in between the three resonance forms and is difficult to draw satisfactorily.

Naphthalene and other polycyclic hydrocarbons show many of the chemical properties we associate with aromaticity. Thus, measurements of heats of hydrogenation show an aromatic stabilization energy approximately 60 kcal/mol (250 kJ/mol) greater than might be expected if naphthalene had five isolated double bonds. Furthermore, naphthalene reacts slowly with electrophilic reagents such as bromine to give substitution products rather than double-bond addition products.



How can we explain the aromaticity of naphthalene? The orbital picture of naphthalene in Figure 15.16 shows a fully conjugated cyclic pi electron system, with *p* orbital overlap both around the 10-carbon periphery of the molecule and across the central bond. Since 10 pi electrons is a Hückel

number, there is a high degree of pi electron delocalization and consequent aromaticity in naphthalene.



**Figure 15.16** An orbital picture of naphthalene, showing that the 10 pi electrons are fully delocalized throughout both rings.

PROBLEM

- 15.16 Look at the three resonance structures of naphthalene, and then account for the fact that not all carbon–carbon bonds have the same length. The C1–C2 bond is 1.36 Å long, whereas the C2–C3 bond is 1.39 Å long.

PROBLEM

- 15.17 Naphthalene is sometimes represented with circles in each ring to represent aromaticity:



The difficulty with this representation is that it's not immediately apparent how many pi electrons are present. How many pi electrons are in each circle?

## 15.12 Spectroscopy of Aromatic Compounds

### INFRARED SPECTROSCOPY

Aromatic rings can be readily detected by infrared spectroscopy since they show a characteristic C–H stretching absorption at  $3030\text{ cm}^{-1}$  and a characteristic series of peaks in the  $1450\text{--}1600\text{ cm}^{-1}$  range. The aryl C–H band at  $3030\text{ cm}^{-1}$  is generally of rather low intensity and occurs just to the left of a normal saturated C–H band. As many as four aryl absorptions are

sometimes observed in the  $1450\text{--}1600\text{ cm}^{-1}$  region owing to complex molecular motions of the ring itself. Two bands, one at  $1500\text{ cm}^{-1}$  and one at  $1600\text{ cm}^{-1}$ , are usually the most intense. In addition, aromatic compounds show strong absorptions in the  $690\text{--}900\text{ cm}^{-1}$  range due to C—H out-of-plane bending. The exact position of these absorptions is diagnostic of the substitution pattern of the aromatic ring:

Monosubstituted:	$690\text{--}710\text{ cm}^{-1}$	<i>m</i> -Disubstituted:	$690\text{--}710\text{ cm}^{-1}$
	$730\text{--}770\text{ cm}^{-1}$		$810\text{--}850\text{ cm}^{-1}$
<i>o</i> -Disubstituted:	$735\text{--}770\text{ cm}^{-1}$	<i>p</i> -Disubstituted:	$810\text{--}840\text{ cm}^{-1}$

The infrared spectrum of toluene in Figure 15.17 shows these characteristic absorptions.

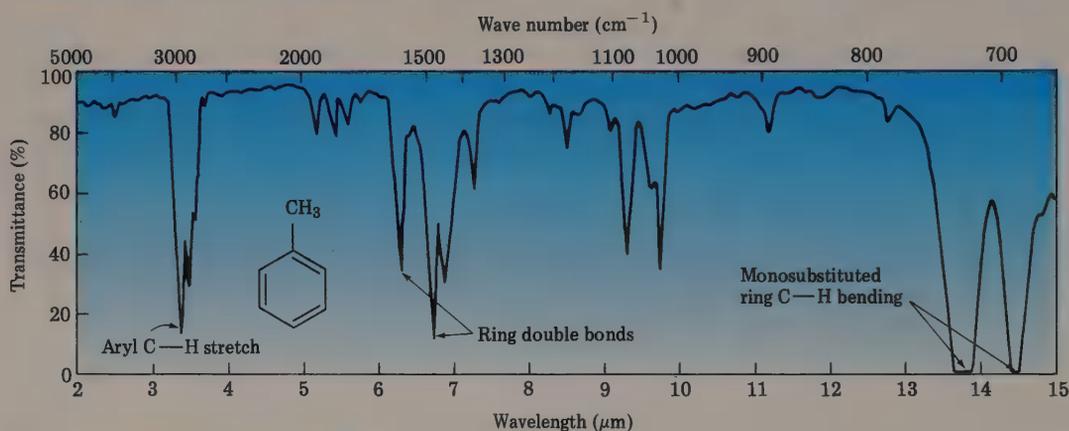


Figure 15.17 The infrared spectrum of toluene.

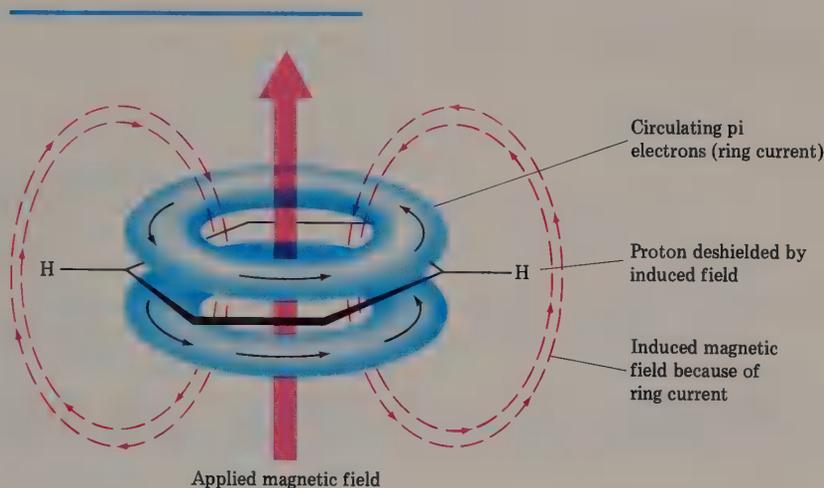
## ULTRAVIOLET SPECTROSCOPY

Aromatic rings are also detectable by ultraviolet spectroscopy because they contain a conjugated pi electron system. In general, aromatic compounds show a series of bands, with a fairly intense absorption near 205 nm and a less intense absorption in the 255–275 nm range. The presence of these bands in the ultraviolet spectrum of a molecule of unknown structure is a sure indication that an aromatic ring is present.

## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

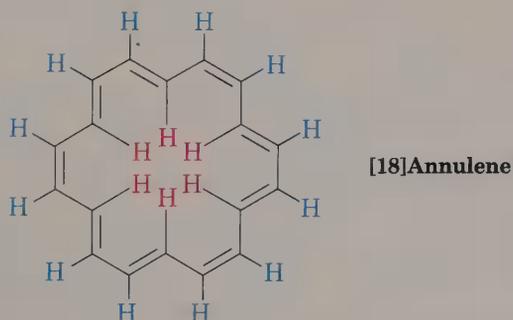
Hydrogens directly bonded to an aromatic ring are easily identifiable in the  $^1\text{H}$  NMR spectrum. These hydrogens are strongly deshielded by the aromatic ring and absorb between 6.5 and 8.0 ppm downfield from the TMS standard. The spins of nonequivalent aryl protons on substituted rings often couple with each other, giving rise to spin–spin splitting patterns that can give information about the substitution pattern of the ring.

Much of the difference in chemical shift between aryl protons (6.5–8.0  $\delta$ ) and vinylic protons (4.5–6.5  $\delta$ ) is due to a special property of aromatic rings called **ring current**. When an aromatic ring is oriented perpendicularly to a strong magnetic field, the pi electrons circulate around the ring and induce a tiny local magnetic field. This induced field opposes the applied field in the middle of the ring but *reinforces* the applied field outside the ring (Figure 15.18). Aryl protons are therefore deshielded; they experience an effective magnetic field greater than the applied field and thus come into resonance at a lower applied field.



**Figure 15.18** The origin of aromatic ring current. Aryl protons are deshielded by the induced magnetic field caused by pi electrons circulating in the molecular orbitals of the aromatic ring.

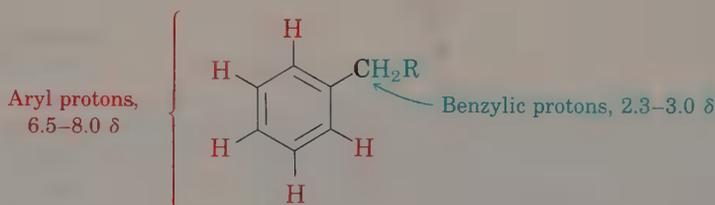
Note that the existence of an aromatic ring current predicts different effects inside and outside the ring. If an aromatic ring were large enough to have both “inside” and “outside” protons, those protons on the outside should be deshielded and absorb at a field lower than normal, but those protons on the inside should be *shielded* and absorb at a field higher than normal. This prediction has been strikingly borne out by studies on [18]annulene, an 18-pi-electron cyclic conjugated polyene that contains a Hückel number of electrons ( $4n + 2$ , where  $n = 4$ ) and is large enough to have both inside and outside protons:



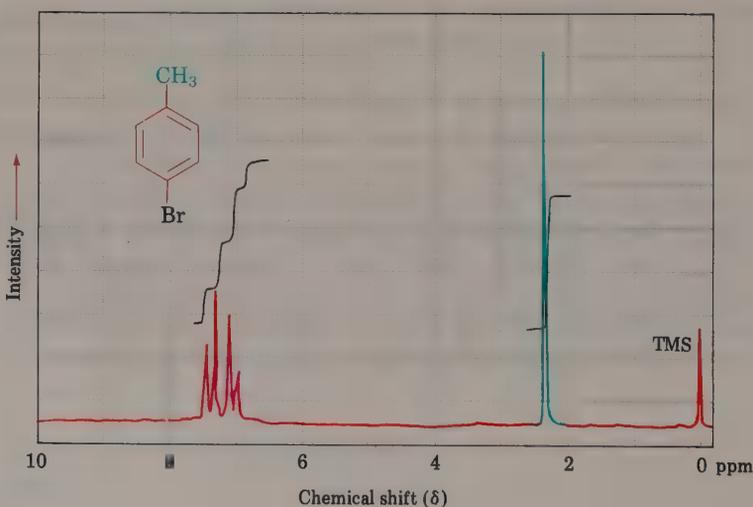
The  $^1\text{H}$  NMR spectrum of [18]annulene is just as predicted: The 6 inside protons are strongly shielded by the aromatic ring current and absorb at 3.0 ppm *upfield* from TMS, while the 12 outside protons are strongly deshielded and absorb in the typical aryl region at 9.3 ppm downfield from TMS.

The presence of a ring current is characteristic of all Hückel aromatic molecules and serves as an excellent test of aromaticity. For example, benzene, a six-pi-electron aromatic molecule, absorbs at 7.37  $\delta$ , but cyclooctatetraene, an eight-pi-electron nonaromatic molecule, absorbs at 5.78  $\delta$ .

Protons on carbon next to aromatic rings also show distinctive absorptions in the NMR spectrum. Benzylic protons normally absorb downfield from other alkane protons in the region from 2.3 to 3.0  $\delta$ .



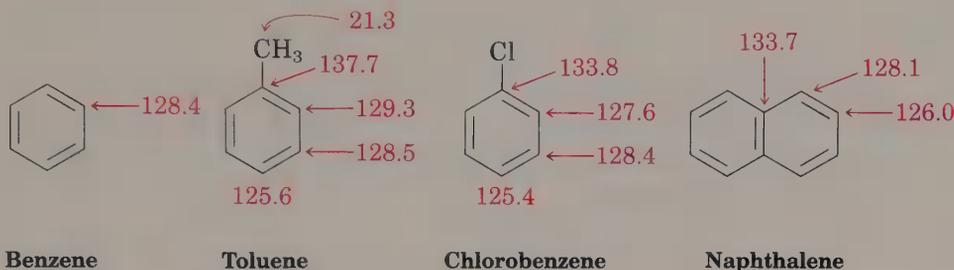
The  $^1\text{H}$  NMR spectrum of *p*-bromotoluene shown in Figure 15.19 displays some of the features just discussed. The aryl protons absorb in a quartet from 6.9 to 7.6  $\delta$ , and the benzylic methyl protons absorb as a sharp singlet at 2.3  $\delta$ . Integration of the spectrum reveals the expected 4:3 ratio of peak areas.



**Figure 15.19** The  $^1\text{H}$  NMR spectrum of *p*-bromotoluene.

Carbon atoms of an aromatic ring absorb in the range 110–160  $\delta$  in the  $^{13}\text{C}$  NMR spectrum, as indicated by the examples in Figure 15.20. These resonances are easily distinguished from those of alkane carbons but occur

in the same range as alkene carbons. Thus, the presence of  $^{13}\text{C}$  absorptions at 110–160  $\delta$  does not in itself establish the presence of an aromatic ring. Confirming evidence from infrared, ultraviolet, or  $^1\text{H}$  NMR is needed.



**Figure 15.20** Some  $^{13}\text{C}$  NMR absorptions of aromatic compounds ( $\delta$  units).

A summary of the kinds of information obtainable from different spectroscopic techniques is given in Table 15.3.

**TABLE 15.3** Summary of Spectroscopic Information on Aromatic Compounds

<i>Kind of spectroscopy</i>	<i>Absorption position</i>	<i>Interpretation</i>
Infrared ( $\text{cm}^{-1}$ )	3030	Aryl C–H stretch
	1500 and 1600	Two intense absorptions due to ring motions
	690–900	Intense C–H out-of-plane bending
Ultraviolet (nm)	205	Intense absorption
	260	Weak absorption
$^1\text{H}$ NMR ( $\delta$ )	2.3–3.0	Benzylic protons
	6.5–8.0	Aryl protons
$^{13}\text{C}$ NMR ( $\delta$ )	110–160	Aromatic ring carbons

## 15.13 Summary and Key Words

The term **aromatic** is used for historical reasons to refer to the class of compounds related structurally to benzene. Many medicinally useful compounds contain aromatic, benzene-like parts.

Aromatic compounds are systematically named according to IUPAC rules, but many common names are also used. Disubstituted benzenes are named as **ortho** (1,2-disubstituted), **meta** (1,3-disubstituted), or **para** (1,4-disubstituted) derivatives. The  $-\text{C}_6\text{H}_5$  unit itself is referred to as a **phenyl** group, and the  $-\text{CH}_2\text{C}_6\text{H}_5$  unit is a **benzyl** group.

Benzene is described by resonance theory as a resonance hybrid of two equivalent Kekulé structures. Neither structure is correct by itself—the true structure of benzene is intermediate between the two:



Benzene is described by molecular orbital theory as a cyclic conjugated molecule with six pi electrons. According to the **Hückel rule**, a cyclic conjugated molecule must have  $4n + 2$  pi electrons, where  $n = 0, 1, 2, 3,$  and so on, in order to be aromatic.

Benzene has five key characteristics:

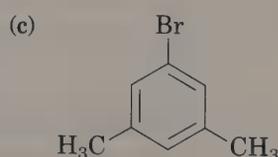
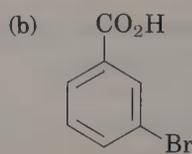
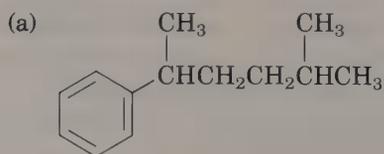
1. Benzene is a cyclic, planar, conjugated molecule.
2. Benzene is unusually stable. Its heat of hydrogenation is approximately 36 kcal/mol less than we might expect for a cyclic triene.
3. Benzene reacts slowly with electrophiles to give substitution products in which cyclic conjugation is retained, rather than alkene addition products in which conjugation is destroyed.
4. Benzene is symmetrical. All carbon–carbon bonds are equivalent and have a length of 1.39 Å, a value that is intermediate between normal single- and double-bond lengths.
5. Benzene has a Hückel number of pi electrons,  $4n + 2$ , where  $n = 1$ . The pi electrons are delocalized over all six carbons, and there is a doughnut-shaped ring of electron density above and below the plane of the ring.

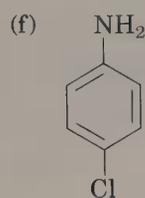
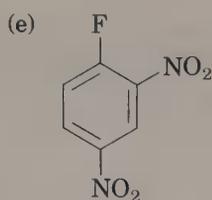
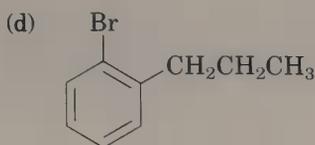
Other kinds of molecules besides benzenoid compounds can also be aromatic according to the Hückel  $4n + 2$  definition. For example, the cyclopentadienyl anion and the cycloheptatrienyl cation are both aromatic ions. **Heterocyclic compounds**, which have atoms other than carbon in the ring, can also be aromatic. Pyridine, a six-membered nitrogen-containing heterocycle, resembles benzene electronically, whereas pyrrole, a five-membered heterocycle, resembles the cyclopentadienyl anion.

All of the spectroscopic techniques studied previously are applicable to the structure elucidation of aromatic compounds. Infrared, ultraviolet, and NMR spectroscopies all show characteristic aromatic absorption peaks.

### ADDITIONAL PROBLEMS .....

15.18 Provide IUPAC names for these compounds:





15.19 Draw structures corresponding to these names:

- (a) 3-Methyl-1,2-benzenediamine      (b) 1,3,5-Benzenetriol  
 (c) 3-Methyl-2-phenylhexane      (d) *o*-Aminobenzoic acid  
 (e) *m*-Bromophenol      (f) 2,4,6-Trinitrophenol (picric acid)  
 (g) *p*-Iodonitrobenzene

15.20 Draw and name all possible isomeric

- (a) Dinitrobenzenes      (b) Bromodimethylbenzenes      (c) Trinitrophenols

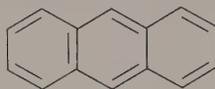
15.21 Draw and name all possible aromatic compounds with the formula  $C_7H_7Cl$ .

15.22 Draw and name all possible aromatic compounds with the formula  $C_8H_9Br$ . (There are 14.)

15.23 Propose structures for aromatic hydrocarbons that meet these descriptions:

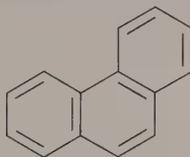
- (a)  $C_9H_{12}$ ; gives only one  $C_9H_{11}Br$  product on aromatic substitution by bromine  
 (b)  $C_{10}H_{14}$ ; gives only one  $C_{10}H_{13}Cl$  product on aromatic substitution by chlorine  
 (c)  $C_8H_{10}$ ; gives three  $C_8H_9Br$  products on aromatic substitution by bromine  
 (d)  $C_{10}H_{14}$ ; gives two  $C_{10}H_{13}Cl$  products on aromatic substitution by chlorine

15.24 There are four resonance structures for anthracene, one of which is shown. Draw the other three.



Anthracene

15.25 There are five resonance structures of phenanthrene, one of which is shown. Draw the other four.



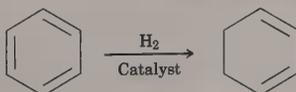
Phenanthrene

15.26 Look at the five resonance structures for phenanthrene (Problem 15.25), and then predict which of its carbon-carbon bonds is shortest.

15.27 Define these terms in your own words:

- (a) Aromaticity      (b) Conjugated  
 (c) Hückel ( $4n + 2$ ) rule      (d) Resonance hybrid

15.28 Table 15.2 gives data on heats of hydrogenation for benzene, 1,3-cyclohexadiene, and cyclohexene. On the basis of these data, calculate the heats of hydrogenation for the partial hydrogenation shown here. Is the reaction exothermic or endothermic?



- 15.29 3-Chlorocyclopropene, on treatment with  $\text{AgBF}_4$ , gives a precipitate of  $\text{AgCl}$  and a stable solution of a species that shows only one absorption in the  $^1\text{H}$  NMR spectrum at  $11.04 \delta$ . How do you explain this result? What is a likely structure for the product, and what is its relation to Hückel's rule?



3-Chlorocyclopropene

- 15.30 Draw an energy diagram for the three molecular orbitals of the cyclopropenyl system (Problem 15.29). How are these three molecular orbitals occupied in the cyclopropenyl anion, cation, and radical? Which of these three is aromatic according to Hückel's rule?
- 15.31 If we were to use the "circle" notation for aromaticity, we would draw the cyclopropenyl cation as shown here. How many pi electrons are represented by the circle in this instance?



- 15.32 Cyclopropanone is a highly reactive molecule because of the large amount of angle strain it contains. Methylcyclopropenone, although more strained than cyclopropanone, is nevertheless quite stable and can even be distilled. How can you account for its stability? [Hint: Consider the polarity of the carbonyl group.]



Cyclopropanone    Methylcyclopropenone

- 15.33 Cycloheptatrienone is a perfectly stable compound, but cyclopentadienone is so reactive that it can't be isolated. What do you suppose accounts for the stability difference between the two?



Cycloheptatrienone    Cyclopentadienone

- 15.34 Which would you expect to be most stable, cyclononatetraenyl radical, cation, or anion?
- 15.35 How might you convert 1,3,5,7-cyclononatetraene to an aromatic substance?
- 15.36 Compound A,  $\text{C}_8\text{H}_{10}$ , yields three monobromo substitution products,  $\text{C}_8\text{H}_9\text{Br}$ , on reaction with bromine. Propose two possible structures. The  $^1\text{H}$  NMR spectrum of A shows a complex four-proton multiplet at  $7.0 \delta$  and a six-proton singlet at  $2.30 \delta$ . What is the correct structure of A?
- 15.37 Azulene,  $\text{C}_{10}\text{H}_8$ , is a beautiful blue hydrocarbon that is isomeric with naphthalene. Unlike naphthalene, however, azulene has a large dipole moment ( $\mu = 1.0 \text{ D}$ ).
- (a) Is azulene a Hückel aromatic compound?

- (b) Draw an orbital picture of azulene.  
 (c) How can you account for the observed dipole moment of azulene?

Azulene ( $\mu = 1.0 \text{ D}$ )

- 15.38 What is the structure of a hydrocarbon that shows a molecular ion at  $m/z = 120$  in the mass spectrum and has the following  $^1\text{H}$  NMR spectrum?

- 7.25  $\delta$ , broad singlet, five protons  
 2.90  $\delta$ , septet,  $J = 7 \text{ Hz}$ , one proton  
 1.22  $\delta$ , doublet,  $J = 7 \text{ Hz}$ , six protons

- 15.39 Propose structures for compounds that fit these descriptions:

(a)  $\text{C}_{10}\text{H}_{14}$

- $^1\text{H}$  NMR: 7.18  $\delta$  (broad singlet, four protons)  
 2.70  $\delta$  (quartet, four protons,  $J = 7 \text{ Hz}$ )  
 1.20  $\delta$  (triplet, six protons,  $J = 7 \text{ Hz}$ )

IR: 745  $\text{cm}^{-1}$

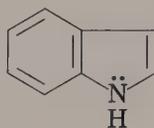
(b)  $\text{C}_{10}\text{H}_{14}$

- $^1\text{H}$  NMR: 7.0  $\delta$  (broad singlet, four protons)  
 2.85  $\delta$  (septet, one proton,  $J = 8 \text{ Hz}$ )  
 2.28  $\delta$  (singlet, three protons)  
 1.20  $\delta$  (doublet, six protons,  $J = 8 \text{ Hz}$ )

IR: 825  $\text{cm}^{-1}$

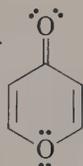
- 15.40 Indole is an aromatic heterocycle that has a benzene ring fused to a pyrrole ring. Draw an orbital picture of indole.

- (a) How many pi electrons are present?  
 (b) What is the electronic relationship of indole to naphthalene?



Indole

- 15.41 On reaction with acid, 4-pyrone is protonated on the carbonyl-group oxygen to give a stable cationic product. Explain the stability of the protonated product.

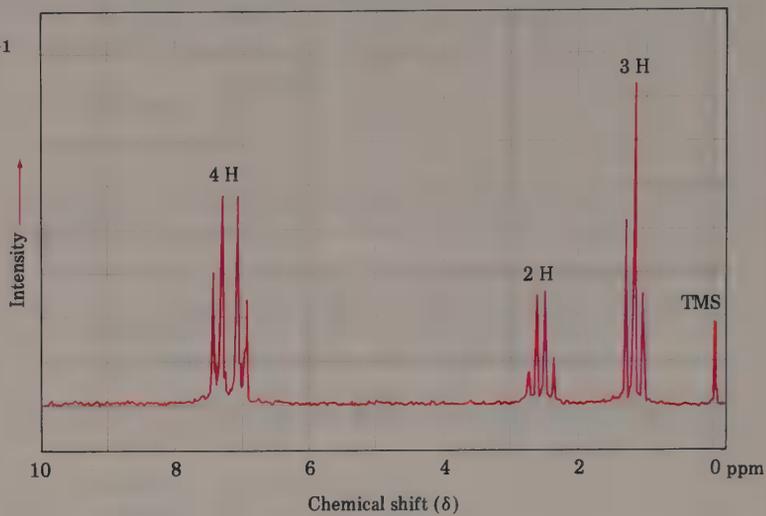


4-Pyrone

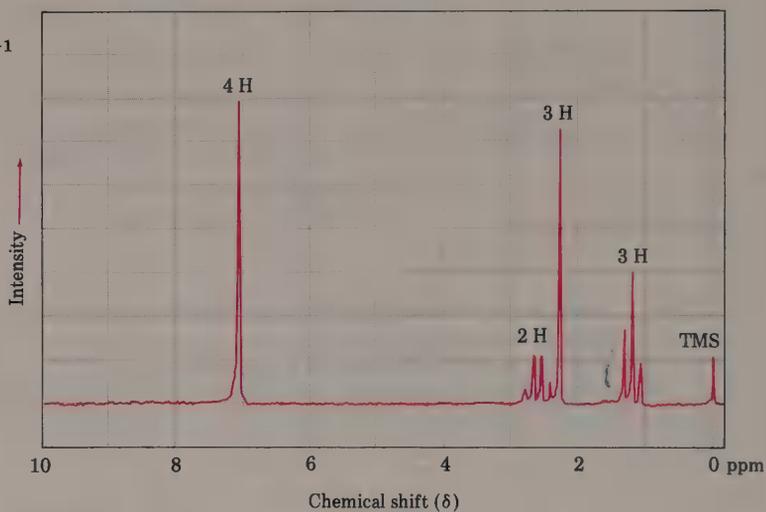
- 15.42 1-Phenyl-2-butene has an ultraviolet absorption at  $\lambda_{\text{max}} = 208 \text{ nm}$  ( $\epsilon = 8000$ ). On treatment with a small amount of strong acid, isomerization occurs and a new substance with  $\lambda_{\text{max}} = 250 \text{ nm}$  ( $\epsilon = 15,800$ ) is formed. Propose a structure for this isomer, and suggest a mechanism for its formation.

15.43 Propose suitable structures for aromatic compounds that have the following  $^1\text{H}$  NMR spectra:

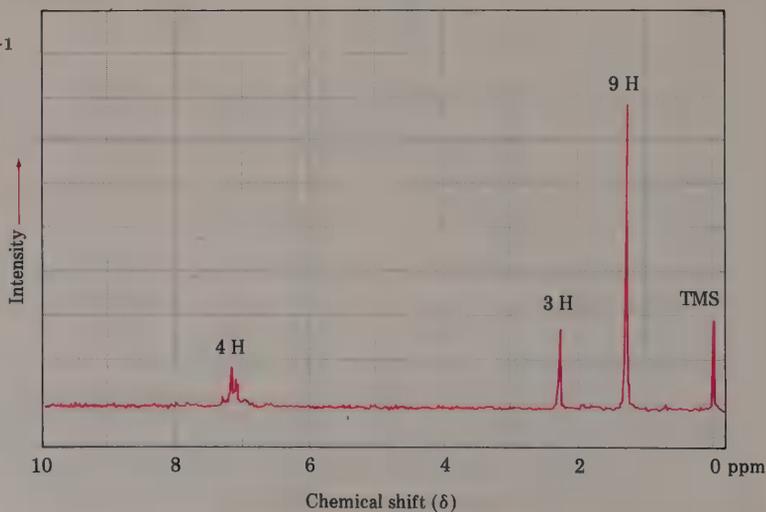
(a)  $\text{C}_8\text{H}_9\text{Br}$   
IR:  $820\text{ cm}^{-1}$



(b)  $\text{C}_9\text{H}_{12}$   
IR:  $750\text{ cm}^{-1}$

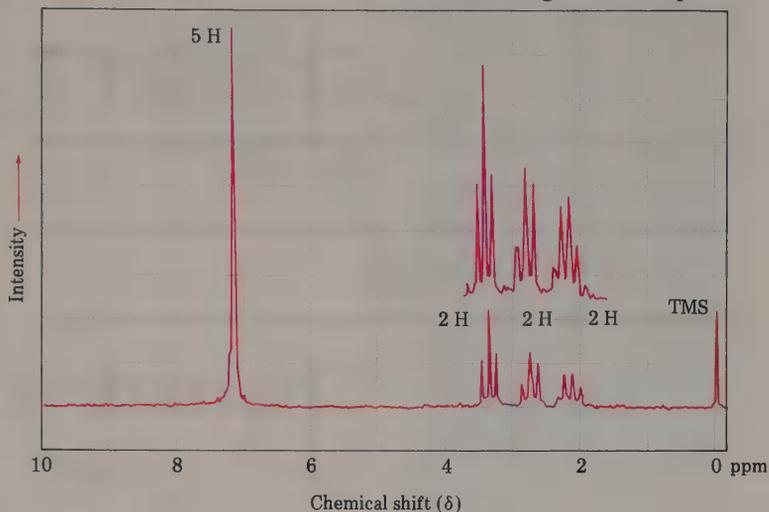


(c)  $\text{C}_{11}\text{H}_{16}$   
IR:  $820\text{ cm}^{-1}$

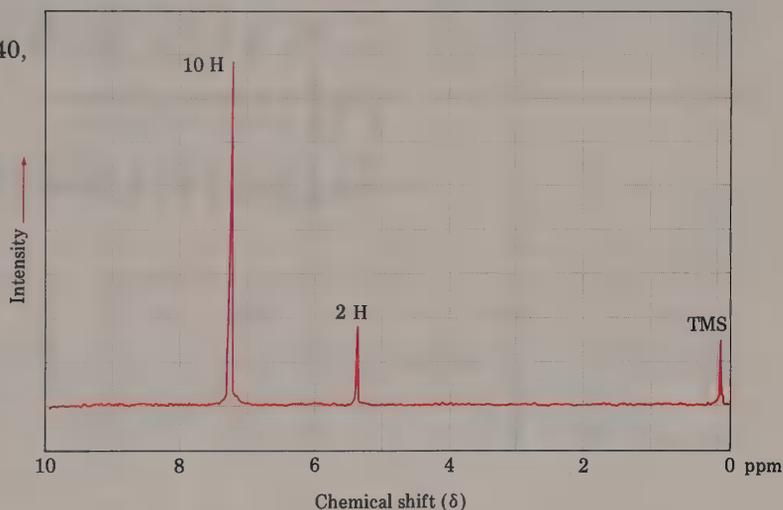


15.44 Propose suitable structures for molecules that have the following  $^1\text{H}$  NMR spectra:

(a)  $\text{C}_9\text{H}_{11}\text{Br}$   
IR: 700,  
760  $\text{cm}^{-1}$



(b)  $\text{C}_{14}\text{H}_{12}$   
IR: 700, 740,  
890  $\text{cm}^{-1}$



15.45 Pentalene is a most elusive molecule that has never been isolated. The pentalene dianion, however, is well known and quite stable. Explain.

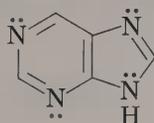


Pentalene



Pentalene dianion

15.46 Purine is a heterocyclic aromatic compound whose derivatives are constituents of DNA and RNA. Why is purine considered to be aromatic? How many  $p$  electrons does each nitrogen donate to the aromatic pi system?



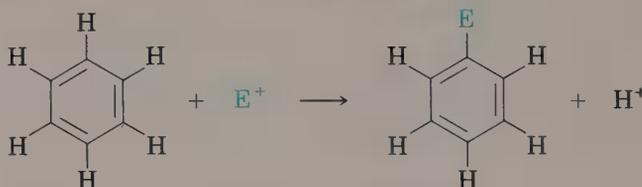
Purine



## Chapter 16

# Chemistry of Benzene: Electrophilic Aromatic Substitution

The most important reaction of aromatic compounds is **electrophilic aromatic substitution**. That is, an electrophile ( $E^+$ ) reacts with an aromatic ring and substitutes for one of the hydrogens.



Many different substituents can be introduced onto the aromatic ring using electrophilic substitution reactions. By choosing the proper conditions and reagents, it's possible to **halogenate** the aromatic ring (substitute a halogen:  $-F$ ,  $-Cl$ ,  $-Br$ , or  $-I$ ), **nitrate** it (substitute a nitro group:  $-\text{NO}_2$ ), **sulfonate** it (substitute a sulfonic acid group:  $-\text{SO}_3\text{H}$ ), **alkylate** it (substitute an alkyl group:  $-\text{R}$ ), or **acylate** it (substitute an acyl group:  $-\text{COR}$ ). Starting from only a few simple materials, we can prepare many thousands of substituted aromatic compounds (Figure 16.1).

All of these reactions (and many more as well) take place by a similar mechanism. Let's begin a study of the process by examining one reaction in detail: the bromination of benzene.

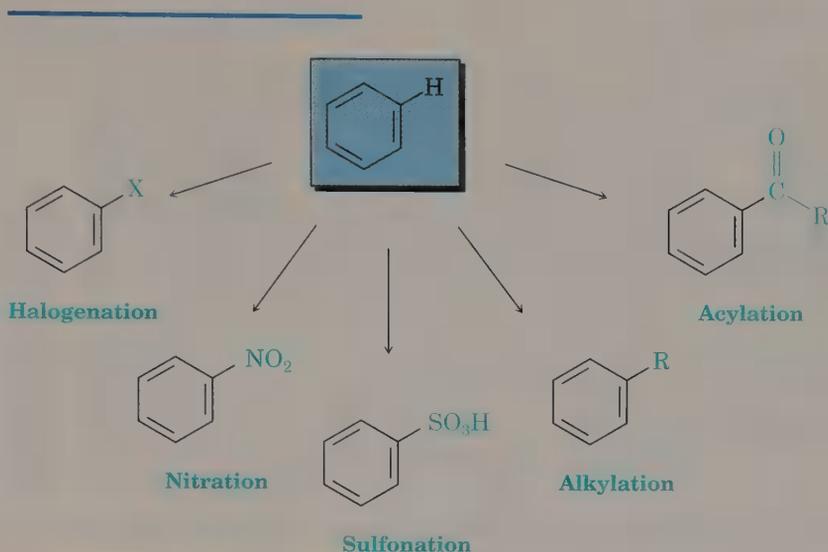
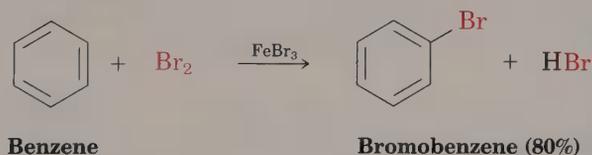


Figure 16.1 Some electrophilic aromatic substitution reactions.

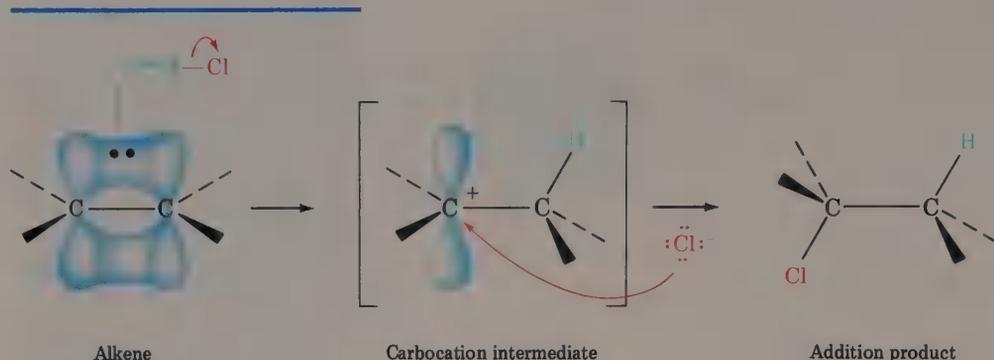
## 16.1 Bromination of Aromatic Rings

A benzene ring, with six pi electrons in a cyclic conjugated system, is a site of electron density. Furthermore, the benzene pi electrons are sterically accessible to attacking reagents because of their location in circular clouds above and below the plane of the ring. Thus, benzene acts as an electron donor (a nucleophile) in most of its chemistry, and most of its reactions take place with electron acceptors (electrophiles). For example, benzene reacts with bromine in the presence of FeBr<sub>3</sub> as catalyst to yield the substitution product, bromobenzene:



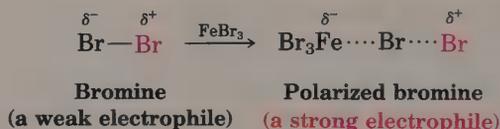
Electrophilic substitution reactions are characteristic of all aromatic rings, not just of benzene and substituted benzenes. Indeed, the ability of a compound to undergo electrophilic substitution is an excellent test of aromaticity.

Before seeing how this electrophilic aromatic *substitution* reaction occurs, let's briefly recall what was said in Chapter 6 about electrophilic *additions* to alkenes. When an electrophile such as HCl adds to an alkene, it approaches the *p* orbitals of the double bond and forms a bond to one carbon, leaving a positive charge at the other carbon. This carbocation intermediate is then attacked by a nucleophile such as chloride ion to yield the addition product (Figure 16.2).

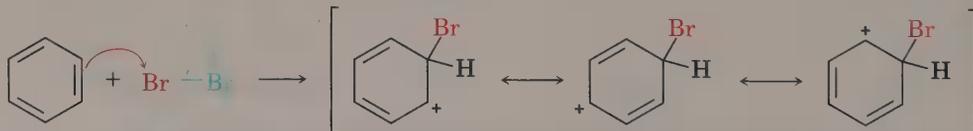


**Figure 16.2** The mechanism of alkene electrophilic addition reactions.

An electrophilic aromatic substitution reaction begins in a similar way, but there are a number of differences. One difference is noticeable immediately: Aromatic rings are much less reactive than alkenes toward electrophiles. For example, bromine in  $\text{CCl}_4$  solution reacts instantly with most alkenes but does not react with benzene. For bromination of benzene to take place, a catalyst such as  $\text{FeBr}_3$  is needed. The catalyst makes the  $\text{Br}_2$  molecule more electrophilic by complexing with it to give an  $\text{FeBr}_4^- \text{Br}^+$  species that reacts as if it were  $\text{Br}^+$ .

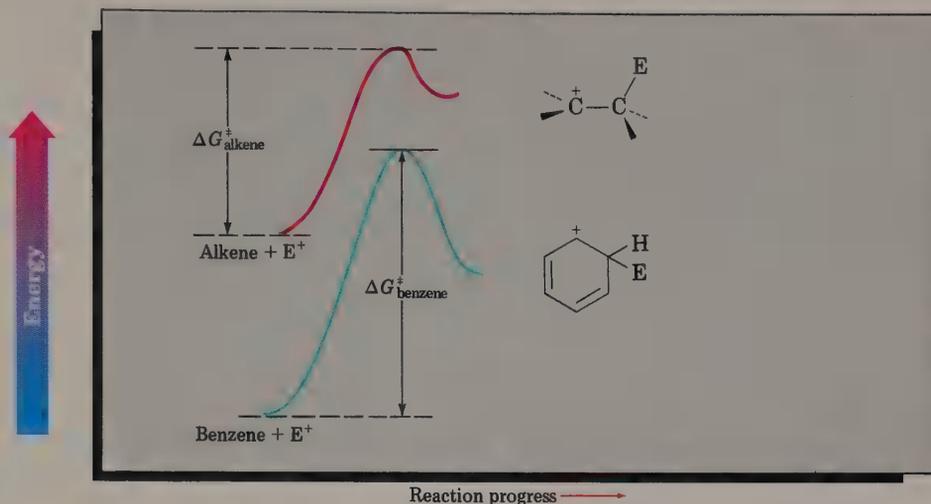


The complexed bromine molecule is then attacked by the pi electron system of the nucleophilic benzene ring in a slow, rate-limiting step to yield a nonaromatic carbocation intermediate. This carbocation is doubly allylic (recall the allyl cation, Section 11.9) and can be written in three resonance forms:



Although stable by comparison with most other carbocations, the intermediate in electrophilic aromatic substitution is nevertheless much less stable than the starting benzene ring with its 36 kcal/mol (152 kJ/mol) of aromatic stability. Thus, electrophilic attack on a benzene ring is endothermic, has a high activation energy, and is therefore a rather slow reaction. Figure 16.3 gives reaction energy diagrams comparing the reaction of an electrophile,  $\text{E}^+$ , with an alkene and with benzene. The benzene reaction is slower (higher  $\Delta G^\ddagger$ ) because the starting material is so much more stable.

A second difference between alkene addition and aromatic substitution occurs after the electrophile has added to the benzene ring to give the car-



**Figure 16.3** A comparison of the reactions of an electrophile with an alkene and with benzene:  $\Delta G^\ddagger$  (alkene)  $\ll$   $\Delta G^\ddagger$  (benzene).

bocation intermediate. Although it would presumably be possible for a nucleophile like bromide ion to attack the carbocation intermediate to yield the addition product dibromocyclohexadiene, this is not observed. Instead, the bromide ion or some other base present in solution abstracts a proton from the bromine-bearing carbon to yield the neutral, aromatic substitution product plus HBr. The net effect of reaction of Br<sub>2</sub> with benzene is the substitution of H<sup>+</sup> by Br<sup>+</sup>. The overall mechanism is shown in Figure 16.4.

Why does the reaction of an electrophile with benzene and other aromatic rings take a different course than its reaction with an alkene? The answer is simple: If *addition* occurred, the 36 kcal/mol stabilization energy of the aromatic ring would be lost and the overall reaction would be endothermic. When *substitution* occurs, though, the stability of the aromatic ring is retained. A reaction energy diagram for the overall process is shown in Figure 16.5.

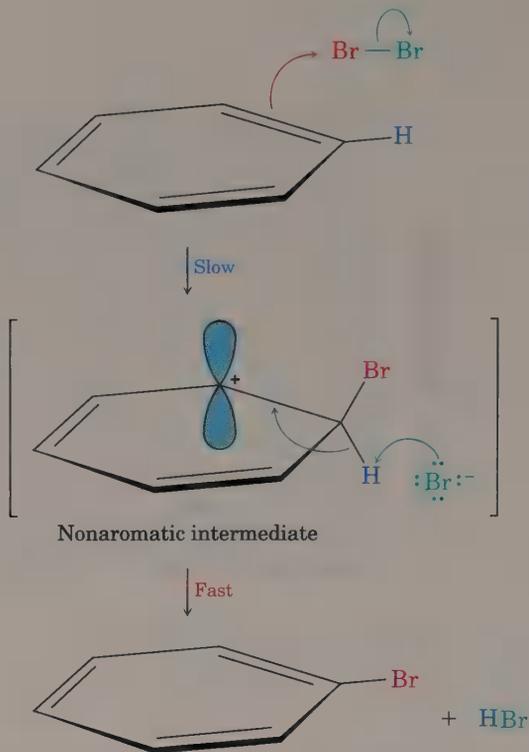
There are many other electrophilic aromatic substitutions besides bromination, and all are thought to occur by the same general mechanism. Let's see briefly what some of these other reactions are.

## 16.2 Other Aromatic Substitutions

### AROMATIC CHLORINATION AND IODINATION

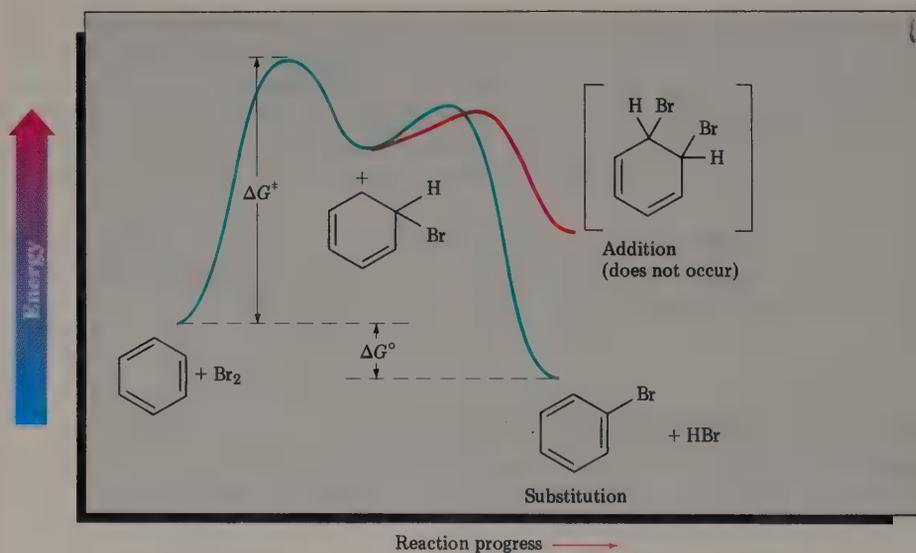
Chlorine and iodine can be introduced into aromatic rings by electrophilic substitution reactions under the proper conditions, but fluorine is too reactive, and poor yields of monofluoroaromatic products are obtained. Aromatic rings react with chlorine in the presence of FeCl<sub>3</sub> catalyst to yield chloro-

An electron pair from the benzene ring attacks  $\text{Br}_2$ , forming a new C-Br bond and leaving a carbocation intermediate.



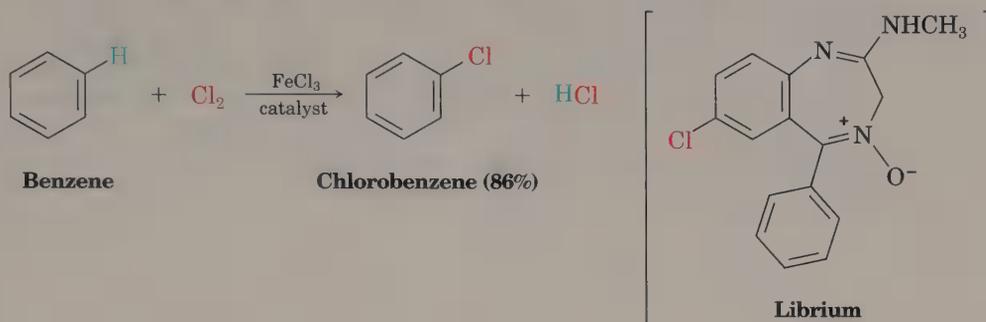
The carbocation intermediate loses a proton and the neutral substitution product forms as two electrons from the C-H bond move into the new aromatic ring.

**Figure 16.4** The mechanism of the electrophilic bromination of benzene.

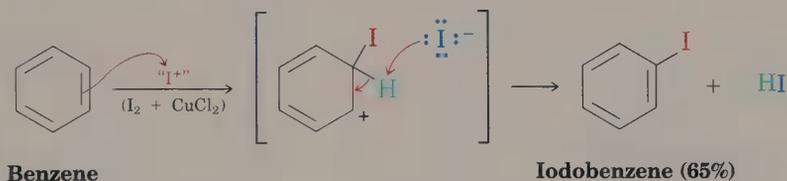


**Figure 16.5** A reaction energy diagram for the electrophilic bromination of benzene.

benzenes. This kind of reaction is used in the synthesis of numerous pharmaceutical agents, including the tranquilizer Librium.



Iodine by itself is unreactive toward aromatic rings, and a promoter is required to obtain a suitable reaction. The best promoters for aromatic iodination are oxidizing agents such as hydrogen peroxide,  $\text{H}_2\text{O}_2$ , or copper salts such as  $\text{CuCl}_2$ . These promoters work by oxidizing molecular iodine to a more powerful electrophilic species that reacts as if it were  $\text{I}^+$ . The aromatic ring then attacks  $\text{I}^+$ , yielding a normal substitution product.



PROBLEM .....

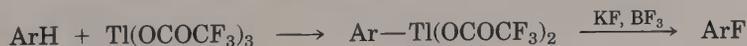
- 16.1 Bromination of toluene gives a mixture of three bromotoluene products. Draw and name them.

PROBLEM .....

- 16.2 Aromatic iodination can be carried out with a number of reagents, including iodine monochloride,  $\text{ICl}$ . What is the direction of polarization of  $\text{ICl}$ ? Propose a mechanism to account for the iodination of an aromatic ring.

PROBLEM .....

- 16.3 Aryl fluorides can be prepared by a two-step process involving electrophilic *thallation* of the aromatic ring (substitution of thallium).



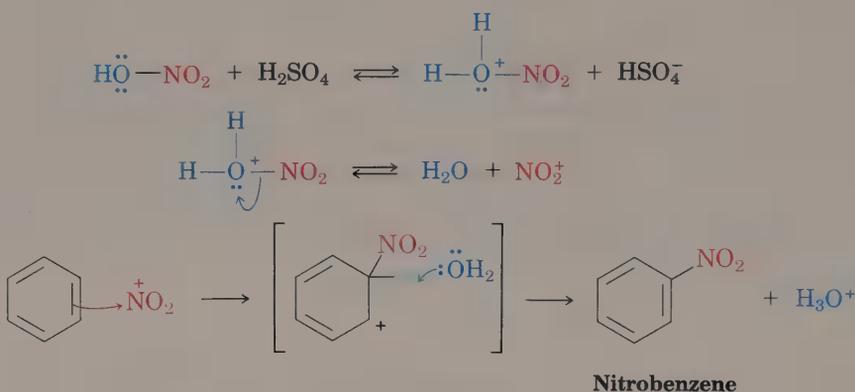
Thallium  
tris(trifluoroacetate)

Propose a mechanism for the thallation of benzene.

.....

## AROMATIC NITRATION

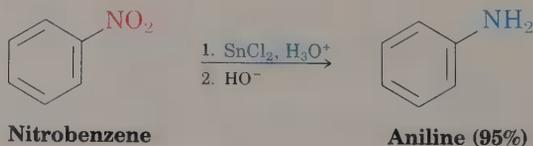
Aromatic rings can be nitrated by reaction with a mixture of concentrated nitric and sulfuric acids. The electrophile in this reaction is the nitronium ion,  $\text{NO}_2^+$ , which is generated from nitric acid by protonation and loss of water. The nitronium ion reacts with benzene to yield a carbocation intermediate in much the same way that we discussed for  $\text{Br}^+$ . Loss of a proton from this intermediate gives the neutral substitution product nitrobenzene (Figure 16.6).



**Figure 16.6** Mechanism of the nitration of benzene to yield nitrobenzene.

When, for some reason, nitration of an aromatic ring with  $\text{HNO}_3/\text{H}_2\text{SO}_4$  is particularly sluggish, the reaction can also be carried out using a pure nitronium salt,  $\text{NO}_2^+ \text{BF}_4^-$ . Nitronium tetrafluoroborate is a white crystalline material that will smoothly nitrate many aromatic compounds at room temperature or below.

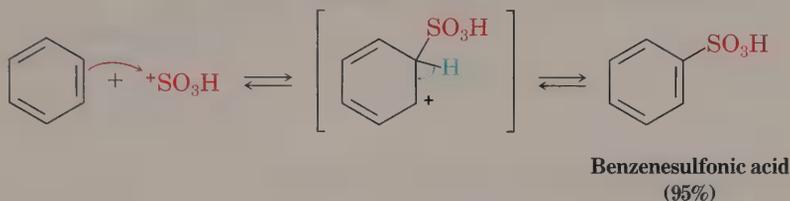
Nitration of aromatic rings is a particularly important reaction because the product nitroarenes can be reduced by reagents such as iron or stannous chloride to yield aminoarenes (substituted anilines). Attachment of a nitrogen to an aromatic ring by the two-step nitration/reduction sequence is a key part of the industrial synthesis of dyes and many pharmaceutical agents. We'll discuss this and other reactions of aromatic nitrogen compounds in Chapter 26.



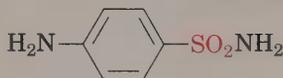
## AROMATIC SULFONATION

Aromatic rings can be sulfonated by reaction with sulfuric acid and sulfur trioxide, a mixture called fuming sulfuric acid ( $\text{H}_2\text{SO}_4 + \text{SO}_3$ ). The reactive

electrophile is either  $\text{HSO}_3^+$  or neutral  $\text{SO}_3$ , depending on reaction conditions. Substitution occurs by the same two-step mechanism seen previously for bromination and nitration. Note, however, that the sulfonation reaction is reversible; it can occur either forward or backward, depending on the reaction conditions. Sulfonation is favored in strong acid, but desulfonation is favored in hot, dilute aqueous acid.

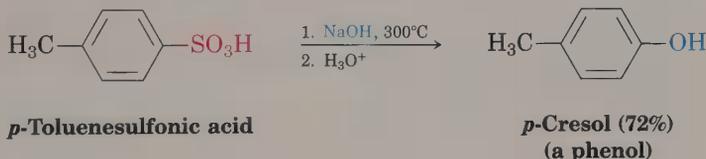


Aromatic sulfonic acids are valuable intermediates in the preparation of dyes and pharmaceuticals. For example, the sulfa drugs such as sulfanilamide were among the first useful antibiotics known. Although largely replaced today by more effective agents, sulfa drugs are still used in the treatment of meningitis. These drugs are prepared commercially by a process that involves aromatic sulfonation as the key step.



Sulfanilamide (an antibiotic)

Aromatic sulfonic acids are also valuable because of the further chemistry they undergo. Heating an arenesulfonic acid with  $\text{NaOH}$  at  $300^\circ\text{C}$  in the absence of solvent yields the corresponding phenol—a net replacement of the sulfonate group by hydroxyl. Yields in this so-called **alkali fusion** process are generally good, but the conditions are so vigorous that the reaction is not compatible with the presence of substituents other than alkyl groups on the aromatic ring.



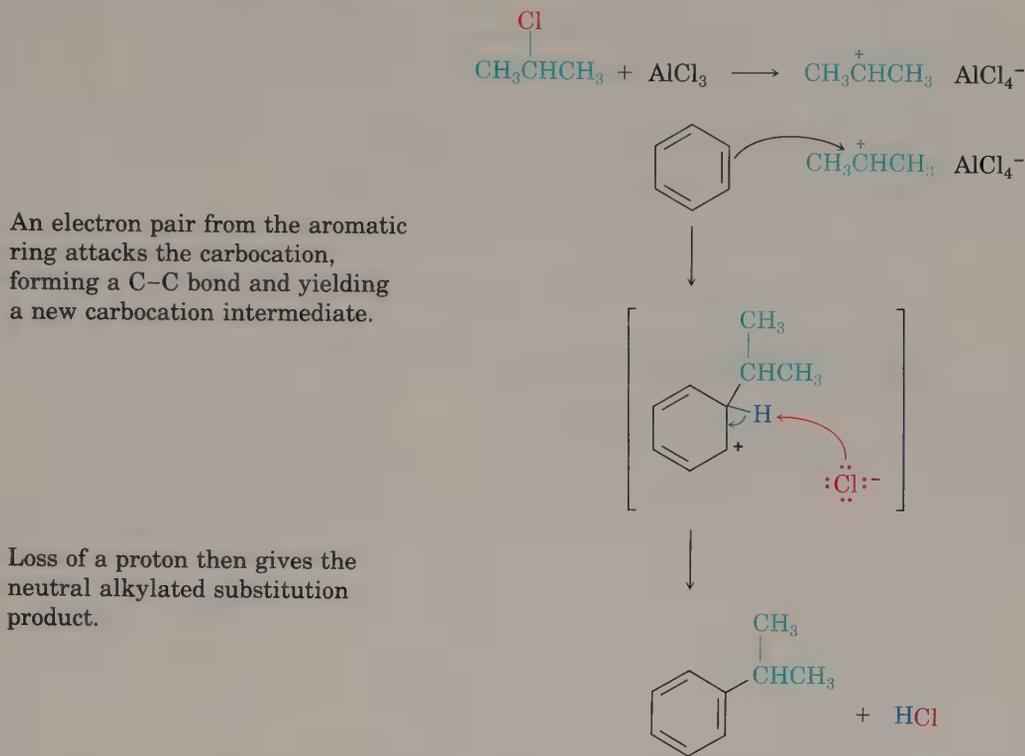
PROBLEM .....

- 16.4 Propose a mechanism for the desulfonation reaction of benzenesulfonic acid to yield benzene. What is the electrophile in this reaction?

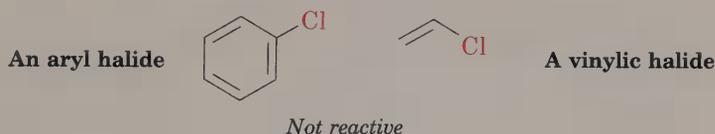
PROBLEM .....

- 16.5 Chlorination of *o*-xylene (dimethylbenzene) yields a mixture of two products, but chlorination of *p*-xylene yields a single product. Explain.



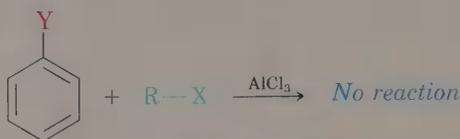


**Figure 16.7** Mechanism of the Friedel–Crafts alkylation reaction.



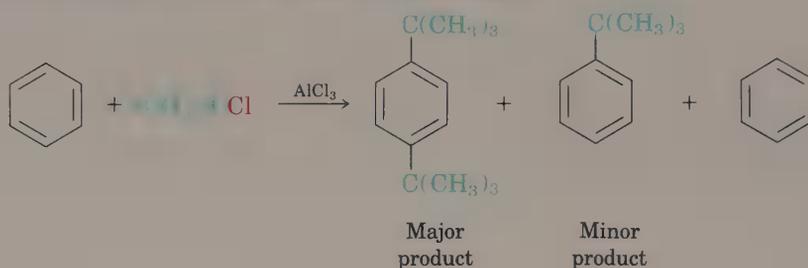
A second limitation is that Friedel–Crafts reactions don't succeed on aromatic rings that are substituted by strongly deactivating groups (Figure 16.8). We'll see in Section 16.5 that the presence of a substituent group on a ring can have a dramatic effect on that ring's subsequent reactivity toward further electrophilic substitution. Rings that contain any of the substituents listed in Figure 16.8 are not reactive enough to attack carbocations.

Yet a third limitation of the Friedel–Crafts alkylation is that it's often difficult to stop the reaction after a single substitution because the alkylated product is often more reactive than the nonalkylated starting material. Once the first group is on the ring, a second substitution reaction is facilitated for reasons we'll discuss in the next section. Thus, we often observe *polyalkylation*. For example, reaction of benzene with 1 mol equiv of 2-chloro-2-methylpropane yields *p*-di-*tert*-butylbenzene as the major product, along with a small amount of *tert*-butylbenzene and unreacted starting material. High yields of monoalkylation product are obtained only when a large excess of benzene is used.

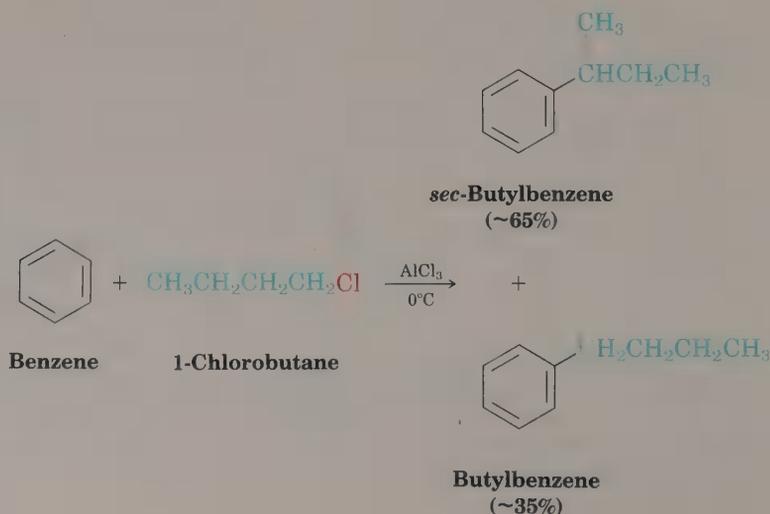


where  $\text{Y} = -\overset{+}{\text{N}}\text{R}_3, -\text{NO}_2, -\text{CN}, -\text{SO}_3\text{H}, -\text{CHO},$   
 $-\text{COCH}_3, -\text{COOH}, -\text{COOCH}_3$   
 $(-\text{NH}_2, -\text{NHR}, -\text{NR}_2)$

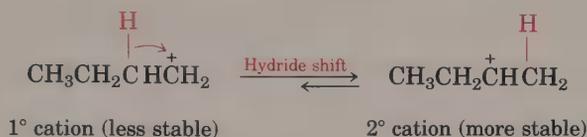
**Figure 16.8** Limitations on the aromatic substrate in Friedel–Crafts reactions.



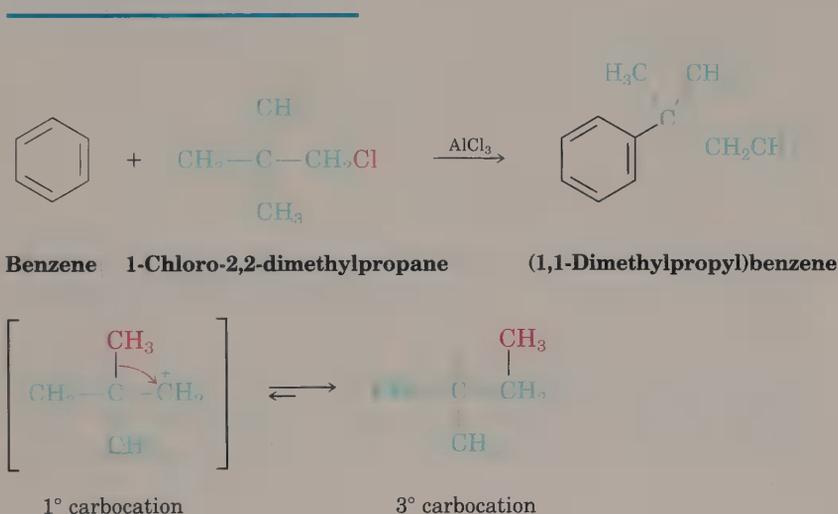
A final limitation to the Friedel–Crafts reaction is that skeletal rearrangements of the alkyl group sometimes occur during reaction, particularly when primary alkyl halides are used. The amount of rearrangement is variable and depends on catalyst, reaction temperature, and even reaction solvent. Less rearrangement is usually found at lower reaction temperatures, but mixtures of products are often obtained. For example, treatment of benzene with 1-chlorobutane gives an approximately 2:1 ratio of rearranged (*sec*-butyl) to unrearranged (*n*-butyl) products when the reaction is carried out at 0°C using  $\text{AlCl}_3$  as catalyst.



These reactions take place by carbocation rearrangements of exactly the same sort we saw earlier during electrophilic additions to alkenes (Section 6.13). For example, the relatively unstable primary butyl carbocation produced by reaction of 1-chlorobutane with  $\text{AlCl}_3$  rearranges to a more stable secondary carbocation by the shift of a hydrogen atom and its electron pair (a hydride ion,  $\text{H}^-$ ) from C2 to C1.



Similarly, carbocation rearrangements can occur by *alkyl* shifts. For example, Friedel–Crafts alkylation of benzene with 1-chloro-2,2-dimethylpropane yields (1,1-dimethylpropyl)benzene as the sole product. The initially formed primary carbocation rearranges to a tertiary carbocation by shift of a methyl group and its electron pair from C2 to C1 (Figure 16.9).



**Figure 16.9** Rearrangement of a primary to a tertiary carbocation during Friedel–Crafts reaction of benzene with 1-chloro-2,2-dimethylpropane.

**PROBLEM** .....

- 16.8** What is the major monosubstitution product that you would expect to obtain from the Friedel–Crafts reaction of benzene with 1-chloro-2-methylpropane in the presence of  $\text{AlCl}_3$ ?

**PROBLEM** .....

- 16.9** The carbocation electrophile in a Friedel–Crafts reaction can be generated in ways other than by reaction of an alkyl chloride with  $\text{AlCl}_3$ . For example, reaction of

benzene with 2-methylpropene in the presence of  $\text{H}_3\text{PO}_4$  yields *tert*-butylbenzene. Formulate a mechanism for this reaction.

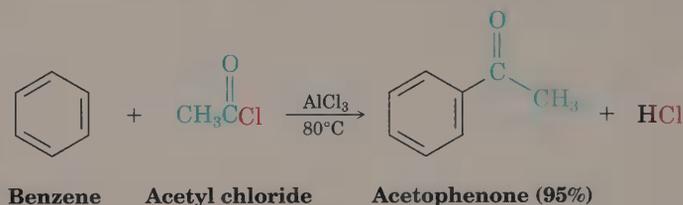
PROBLEM.....

16.10 Which of the following alkyl halides would you expect to undergo Friedel–Crafts reaction *without* rearrangement? Explain.

- (a)  $\text{CH}_3\text{CH}_2\text{Cl}$  (b)  $\text{CH}_3\text{CH}_2\text{CH}(\text{Cl})\text{CH}_3$  (c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$   
 (d)  $(\text{CH}_3)_3\text{CCH}_2\text{Cl}$  (e) Chlorocyclohexane
- .....

## 16.4 Acylation of Aromatic Rings

An **acyl group**,  $-\text{COR}$  (pronounced **a-sil**), is introduced onto the ring when an aromatic compound is allowed to react with a carboxylic acid chloride,  $\text{RCOCl}$ , in the presence of  $\text{AlCl}_3$ . For example, reaction of benzene with acetyl chloride yields the ketone, acetophenone.



The mechanism of **Friedel–Crafts acylation** is similar to that of Friedel–Crafts alkylation. The reactive electrophile is a resonance-stabilized acyl cation, generated by reaction between the acyl chloride and  $\text{AlCl}_3$  (Figure 16.10). As the resonance structures in Figure 16.10 indicate, an acyl cation is stabilized by overlap of the vacant orbital on carbon with a lone-pair orbital of the neighboring oxygen. Once formed, the acyl cation does not rearrange; rather, it is attacked by an aromatic ring to give unarranged substitution product.

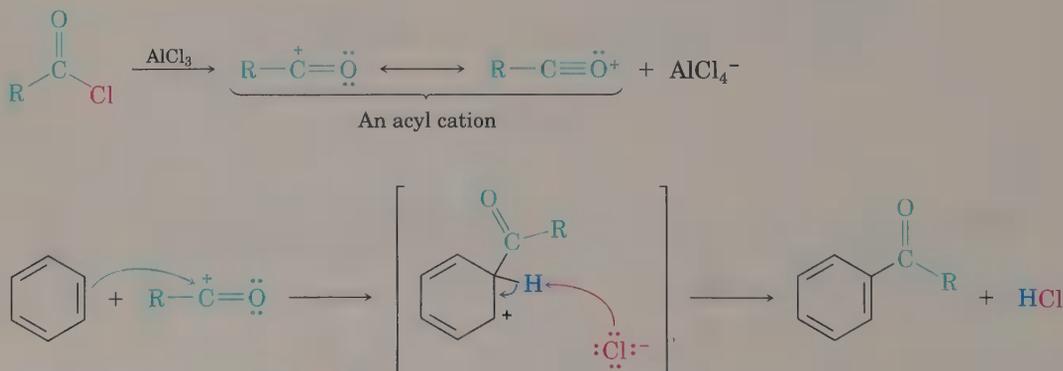


Figure 16.10 Mechanism of the Friedel–Crafts acylation reaction.

Unlike the multiple substitutions that often occur in Friedel–Crafts alkylations, acylations never occur more than once on a ring, because the product acylbenzene is always less reactive than the nonacylated starting material. We'll account for these reactivity differences in the next section.

## 16.5 Reactivity of Aromatic Rings

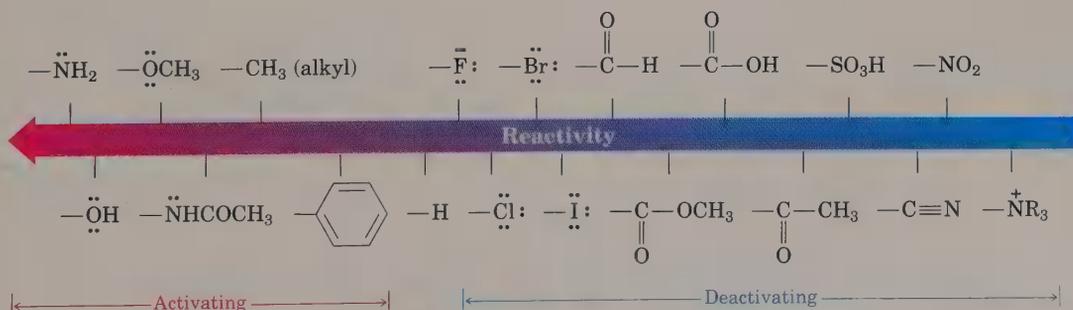
Only one monosubstitution product can result when electrophilic substitution occurs on benzene, but what will happen if we carry out a reaction on a ring that already has a substituent on it? Substituents already present on the ring have two effects:

1. Substituents affect the *reactivity* of the aromatic ring. Some substituents make the ring more reactive than benzene, and some make it less reactive.
2. Substituents affect the *orientation* of the reaction. Three possible disubstituted products can result: ortho, meta, and para. These three products aren't formed in random ratios, though. Instead, the nature of the substituent already present on the benzene ring determines the position of the second substitution.

Let's look at these two effects more closely.

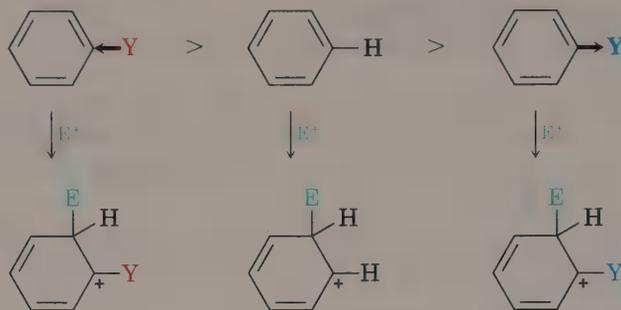
### CLASSIFICATION OF SUBSTITUENTS

Substituents can be classified into two groups: those that *activate* the aromatic ring toward further electrophilic substitution, and those that *deactivate* it. Rings that contain an activating substituent are more reactive than benzene, whereas those with a deactivating substituent are less reactive than benzene. Figure 16.11 lists some groups in both categories.



**Figure 16.11** Activating and deactivating substituents for electrophilic aromatic substitution.

The common feature of all substituents within a category is that all activating groups are able to donate electrons to the ring, thereby stabilizing the carbocation intermediate, while all deactivating groups withdraw electrons from the ring, thereby destabilizing the carbocation intermediate. An aromatic ring with an electron-donating substituent is thus more reactive toward electrophiles, whereas an aromatic ring with an electron-withdrawing substituent is less reactive toward electrophiles.



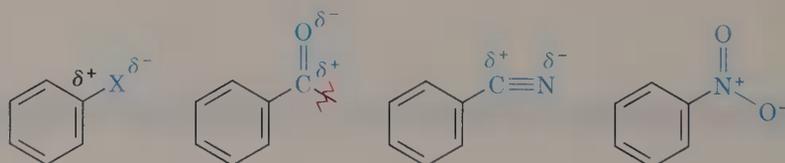
**Y** is an electron donor; carbocation intermediate is more stabilized, and ring is more reactive.

**Y** is an electron acceptor; carbocation intermediate is less stabilized, and ring is less reactive.



## RESONANCE AND INDUCTIVE EFFECTS

A substituent can donate or withdraw electrons from the aromatic ring in two ways: by *inductive effects* and by *resonance effects*. As we've seen previously (Sections 2.4 and 6.11), **inductive effects** are due to the intrinsic electronegativity of atoms and to bond polarity in functional groups. These effects operate by donating or withdrawing electrons through sigma bonds. For example, halogens, carbonyl groups, cyano groups, and nitro groups deactivate an aromatic ring by inductively withdrawing electrons through the sigma bond linking the substituent to the ring.

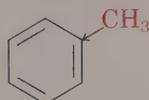


(X = F, Cl, Br, I)

The groups attached to the aromatic rings are inductively electron-withdrawing because of the polarity of their bonds.

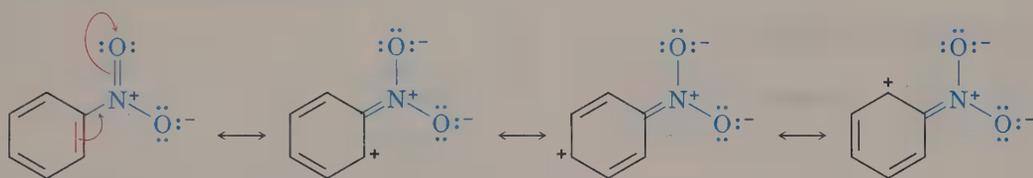
Alkyl groups are inductively electron-donating and therefore activate the ring. The reasons for this aren't fully understood but probably involve

the same factors that cause alkyl substituents to stabilize alkenes (Section 6.7) and carbocations (Section 6.11).

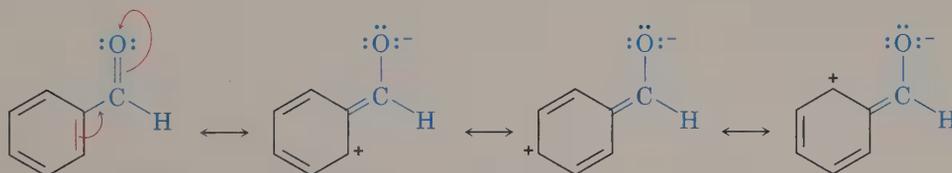


Alkyl group; inductively electron-donating

**Resonance effects** operate by donating or withdrawing electrons through pi bonds by overlap of a *p* orbital on the substituent with a *p* orbital on the aromatic ring. Nitro, cyano, and carbonyl substituents all fit in this category because electrons can flow from the nitrobenzene and benzaldehyde rings onto the substituents, leaving a positive charge in the rings. The carbocation reaction intermediates are therefore destabilized, and the rings are deactivated toward electrophilic attack. Figure 16.12 shows resonance structures illustrating this effect.



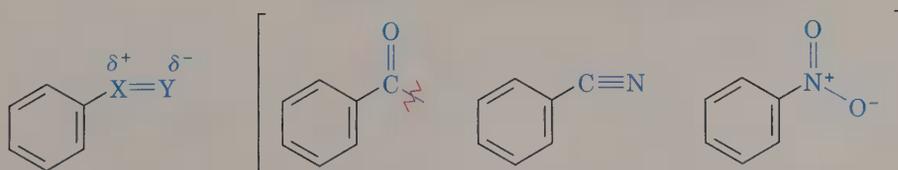
Nitrobenzene



Benzaldehyde

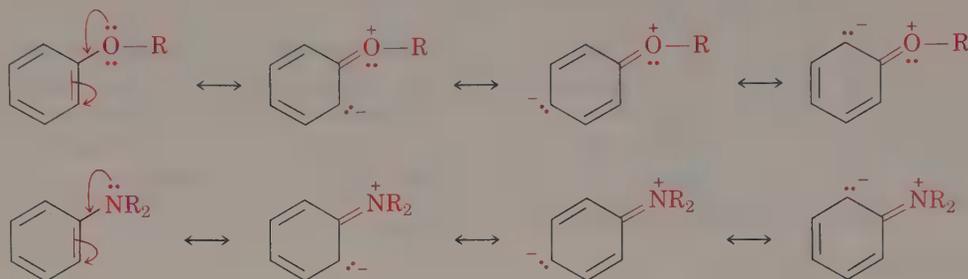
**Figure 16.12** Carbonyl and nitro substituents deactivate aromatic rings by resonance withdrawal of pi electrons. Note that the effect is felt most strongly at ortho and para positions.

Note that substituents with a resonance deactivating effect have the general structure  $-X=Y$ , where the Y atom is more electronegative than X:



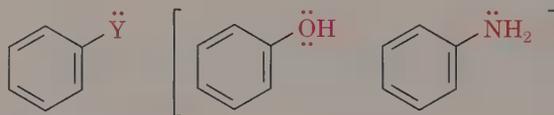
Deactivated rings have this general structure.

Conversely, hydroxyl, methoxyl, and amino substituents *activate* the aromatic ring by resonance effects. Pi electrons flow from the substituents to the ring as shown by the resonance structures in Figure 16.13. This electron donation places a negative charge in the ring, stabilizing the carbocation reaction intermediate and making the ring more reactive toward nucleophiles. (Recall the discussion of conjugation in Section 14.9 where a similar effect was noted.)



**Figure 16.13** Hydroxyl, methoxyl, and amino substituents activate aromatic rings by resonance donation of pi electrons. The effect is felt most strongly at ortho and para positions.

Note that substituents with a resonance activating effect have the general structure  $-\ddot{Y}$ , in which the atom Y has a lone pair of electrons it can donate to the ring:



Activated rings have this general structure.

It probably comes as a surprise to hear that hydroxyl, methoxyl, and amino groups activate an aromatic ring. After all, both oxygen and nitrogen are highly electronegative and would be expected to inductively deactivate a ring in the same way that halogens do. In fact, these groups *do* have an electron-withdrawing inductive effect. We find, however, that the electron-donating resonance effect through pi orbitals is stronger than the electron-withdrawing inductive effect through sigma orbitals. Thus, these substituents have a net activating influence.

**PROBLEM** .....

**16.11** Rank the compounds in each group in the order of their reactivity to electrophilic substitution.

- Nitrobenzene, phenol, toluene, benzene
- Phenol, benzene, chlorobenzene, benzoic acid
- Benzene, bromobenzene, benzaldehyde, aniline

## PROBLEM.....

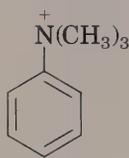
- 16.12 Use Figure 16.11 to account for the fact that Friedel–Crafts alkylations often give polysubstitution but Friedel–Crafts acylations never give polysubstitution.

## PROBLEM.....

- 16.13 Write as many resonance forms as you can for the carbocation intermediates in these reactions. Be sure to consider the resonance effect of the hydroxyl substituent.
- Bromination of phenol at the para position
  - Bromination of phenol at the meta position
  - Bromination of phenol at the ortho position
- At which position(s) does reaction appear most favorable? Which position(s) are least favored?

## PROBLEM.....

- 16.14 Write as many resonance forms as you can for the carbocation intermediates formed by nitration of the *N,N,N*-trimethylanilinium ion at ortho, meta, and para positions. At which position(s) does reaction look more favorable? Explain.



*N,N,N*-Trimethylanilinium ion

## 16.6 Orientation of Reactions on Substituted Aromatic Rings

In addition to affecting the reactivity of an aromatic ring, a substituent can also influence the position of further electrophilic substitution. For example, a methyl substituent shows a strong ortho- and para-directing effect. Nitration of toluene yields predominantly ortho (63%) and para (34%) products, along with only 3% of the meta isomer.

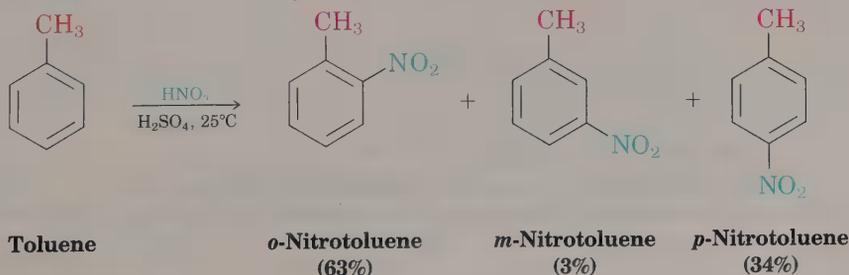
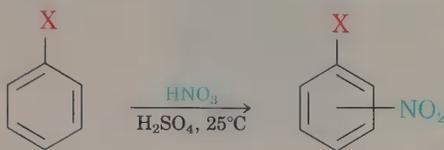




TABLE 16.1 Orientation of Nitration in Substituted Benzenes

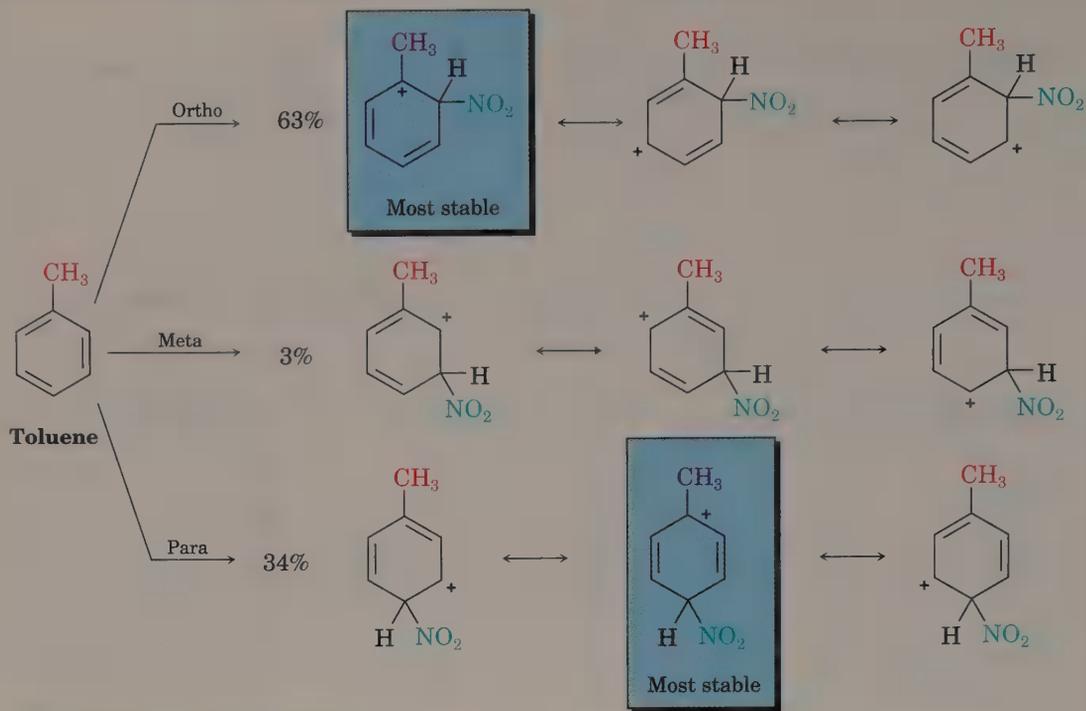


X	Product (%)			X	Product (%)		
	Ortho	Meta	Para		Ortho	Meta	Para
<b>Meta-directing deactivators</b>				<b>Ortho- and para-directing deactivators</b>			
$-\overset{+}{\text{N}}(\text{CH}_3)_3$	2	89	11	$-\text{F}$	13	1	86
$-\text{NO}_2$	7	91	2	$-\text{Cl}$	35	1	64
$-\text{COOH}$	22	77	2	$-\text{Br}$	43	1	56
$-\text{CN}$	17	81	2	$-\text{I}$	45	1	54
$-\text{CO}_2\text{CH}_2\text{CH}_3$	28	66	6	<b>Ortho- and para-directing activators</b>			
$-\text{COCH}_3$	26	72	2	$-\text{CH}_3$	63	3	34
$-\text{CHO}$	19	72	9	$-\ddot{\text{O}}\text{H}$	50	0	50
				$-\ddot{\text{N}}\text{HCOCH}_3$	19	2	79

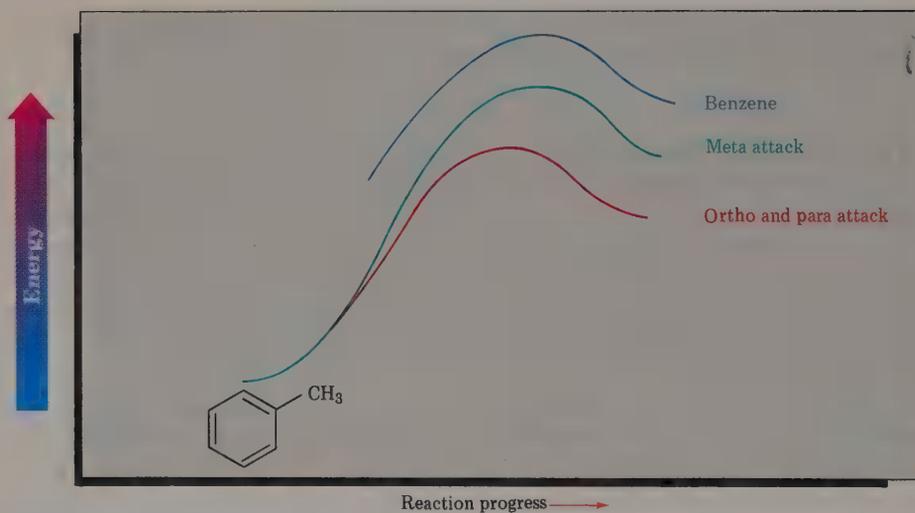
### ORTHO- AND PARA-DIRECTING ACTIVATORS: ALKYL

When groups are classed as ortho- and para-directing activators, it's important to realize that they activate *all* positions on the aromatic ring. These substituents are ortho- and para-directing only because they activate the ortho and para positions more than they activate the meta positions. In other words, substitution in activated benzenes occurs at the position where activation is felt most. Let's look at the nitration of toluene, which is an alkyl-substituted benzene, as an example of how ortho and para directors work (Figure 16.15).

Nitration of toluene might occur either ortho, meta, or para to the methyl group, giving the carbocation intermediates shown in Figure 16.15. Although all three intermediates are resonance stabilized by the remaining two ring double bonds, the ortho and para intermediates are *most* stabilized. For both ortho and para attack, but not for meta, a resonance form places the positive charge directly on the methyl-substituted carbon, where it is in a tertiary position and can best be stabilized by the methyl inductive effect. The ortho and para intermediates are lower in energy than the meta intermediate is and are therefore formed preferentially, as indicated by the reaction energy diagram in Figure 16.16.



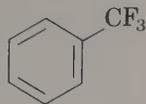
**Figure 16.15** Carbocation intermediates in the nitration of toluene. Ortho and para intermediates are more stable than meta intermediates.



**Figure 16.16** Reaction energy diagram for the nitration of toluene. The carbocation intermediates formed by attack at ortho and para positions are more stable and are therefore formed more rapidly than the intermediate formed by attack at the meta position.

## PROBLEM .....

- 16.16 Which compound would you expect to be more reactive toward electrophilic substitution, toluene or (trifluoromethyl)benzene? Explain.



(Trifluoromethyl)benzene

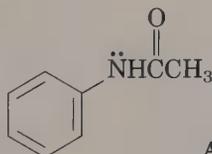
 ORTHO- AND PARA-DIRECTING ACTIVATORS:  
 OH AND NH<sub>2</sub>

Hydroxyl and amino groups and their derivatives are also ortho-para activators, but for a different reason than for alkyl groups. Hydroxyl and amino groups exert their activating influence through a strong resonance electron-donating effect, which is most pronounced at the ortho and para positions. The resonance forms of phenol shown previously in Figure 16.13 indicate how the oxygen lone-pair electrons can be shared only in the ortho and para positions, but not in the meta position.

When phenol is nitrated, the results shown in Figure 16.17 are obtained. Although three products are possible, only ortho and para attack is observed. All three of the possible carbocation intermediates are stabilized by resonance, but the intermediates from ortho and para attack are stabilized most. Only in ortho and para attack are there resonance forms in which the positive charge is stabilized by donation of an electron pair from oxygen. The product of meta attack has no such stabilization.

## PROBLEM .....

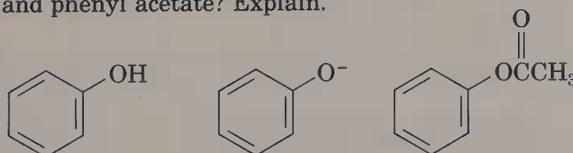
- 16.17 Acetanilide is much less reactive than aniline toward electrophilic substitution. Explain.



Acetanilide

## PROBLEM .....

- 16.18 What reactivity order toward electrophilic substitution would you expect for phenol, phenoxide ion, and phenyl acetate? Explain.

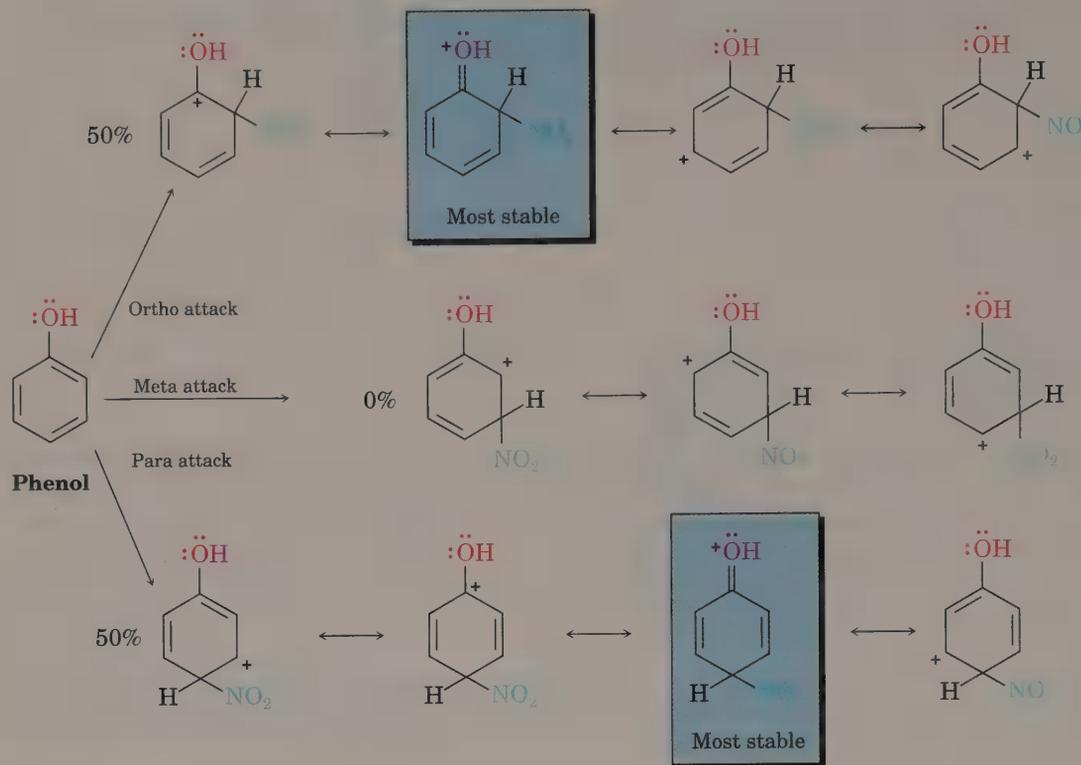


Phenol

Phenoxide ion

Phenyl acetate

## .....

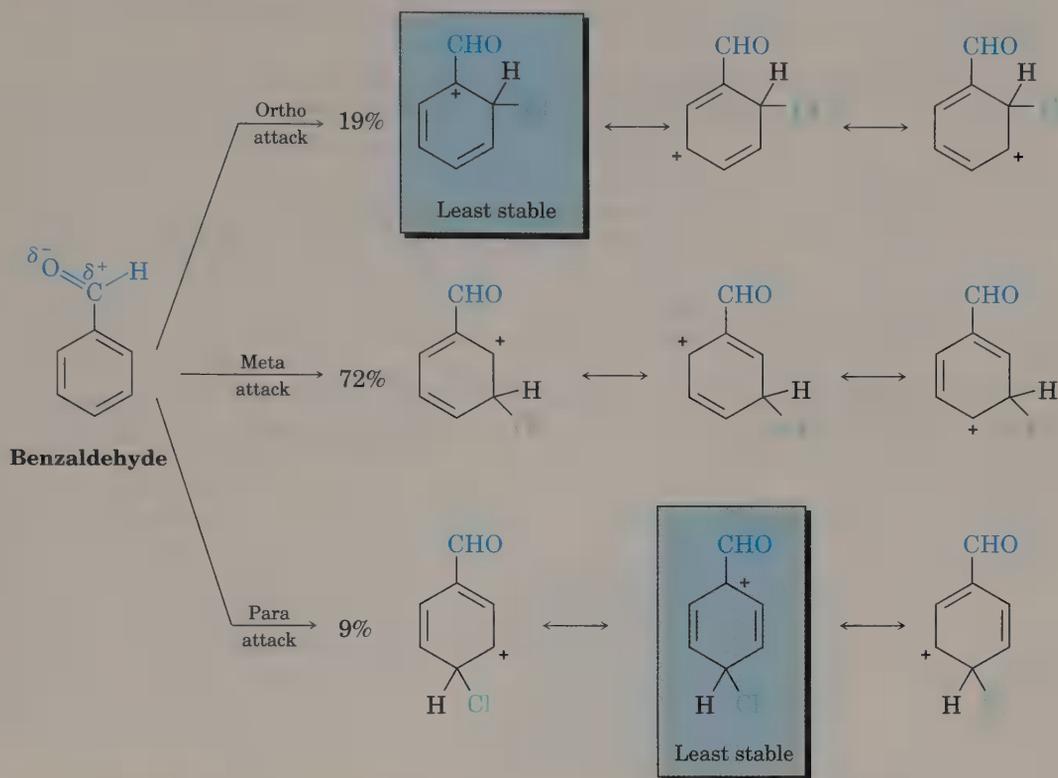


**Figure 16.17** Carbocation intermediates in the nitration of phenol. The ortho and para intermediates are more stable than the meta intermediate because of resonance donation of the oxygen electrons.

## META-DIRECTING DEACTIVATORS

When groups such as carbonyl, nitro, and cyano are classified as deactivators, it's important to realize that these groups deactivate *all* positions on the ring. Meta directors exert their directing influence because they deactivate the meta position less than they deactivate the ortho and para positions. In other words, substitution in deactivated benzenes occurs at the least deactivated position.

We can explain the influence of meta-directing deactivators by the same kinds of arguments just used for ortho- and para-directing activators. For example, let's look at the chlorination of benzaldehyde, shown in Figure 16.18. Of the three possible intermediates, the carbocations produced by reaction at ortho and para positions are least stable, and the cation produced by reaction at the meta position is therefore most stable.



**Figure 16.18** Intermediates in the chlorination of benzaldehyde. The meta intermediate is more stable than the ortho or para intermediate.

Meta-directing deactivators such as carbonyl, nitro, and cyano exert their influence through a combination of inductive and resonance effects that reinforce each other. Inductively, both ortho and para intermediates are destabilized because the positive charge of the carbocationic intermediate can be placed directly on the ring carbon atom that bears the deactivating group, where it is disfavored by a repulsive interaction with the positively polarized carbon atom (Figure 16.18). The meta intermediate is therefore the most favored.

Resonance electron withdrawal is also felt at the ortho and para positions, rather than at the meta position as the dipolar resonance forms of benzaldehyde shown previously in Figure 16.12 indicate. Reaction with an electrophile therefore occurs at the meta position.

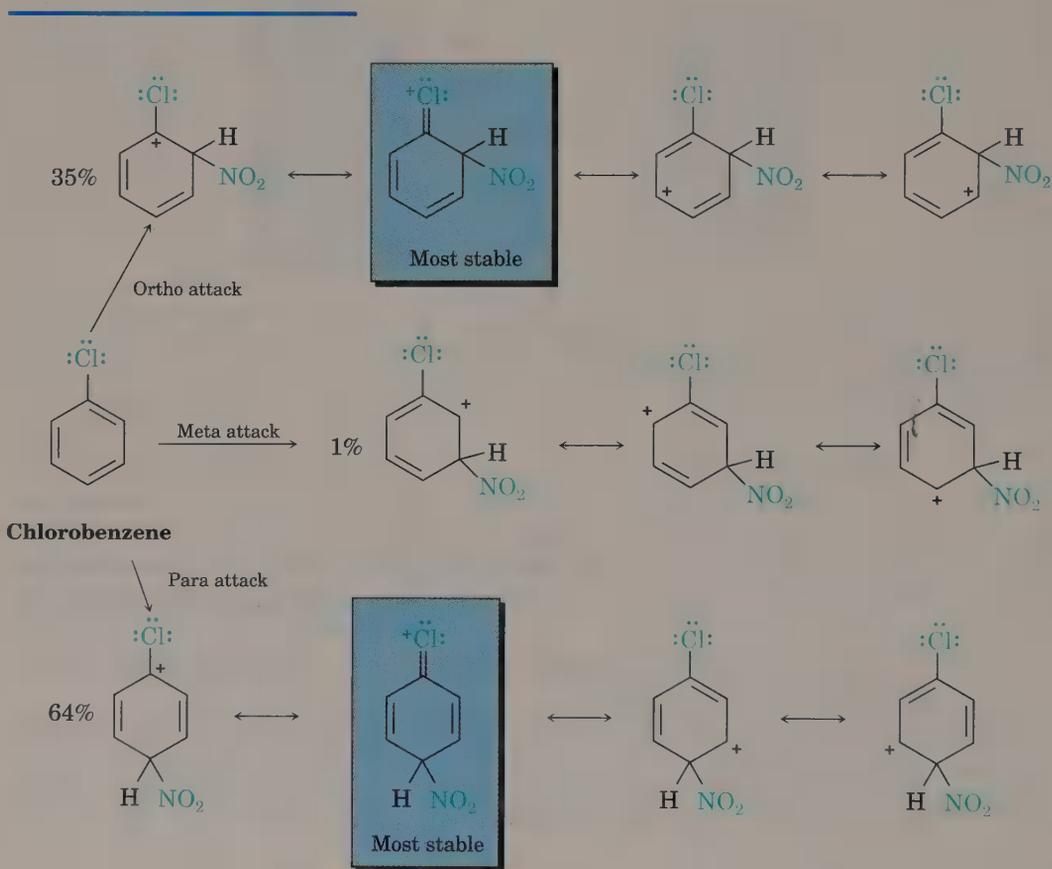
**PROBLEM** .....

- 16.19** Draw resonance structures for the intermediates from attack of an electrophile at the ortho, meta, and para positions of nitrobenzene. Which intermediates are most favored?
- .....

## ORTHO- AND PARA-DIRECTING DEACTIVATORS: HALOGENS

Halogen substituents occupy a unique position since they are deactivating yet have an ortho- and para-directing effect. Why should this be?

As we saw in the previous section, halogen substituents are deactivating because of their strong electron-withdrawing inductive effect. Unlike other deactivating groups, though, halogens deactivate the ortho and para positions *less* than they deactivate the meta position. The reasons for this behavior can be seen by looking at the three possible intermediates formed on nitration of chlorobenzene (Figure 16.19). Although halogens have an electron-*withdrawing* inductive effect, they have an electron-*donating* resonance effect because of *p* orbital overlap of lone-pair electrons on halogen with the pi orbitals of the aromatic ring. Thus, the halogen substituent can stabilize the positive charge of the carbocation intermediates from ortho



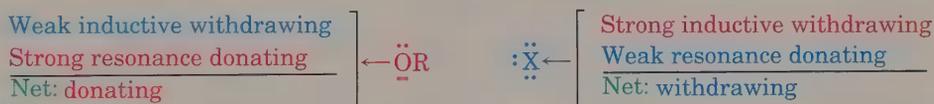
**Figure 16.19** Intermediates in the nitration of chlorobenzene. The ortho and para intermediates are more stable than the meta intermediate because of resonance electron donation of the halogen lone-pair electrons.

and para attack in the same way that hydroxyl and amino substituents can. The meta intermediate, however, has no such stabilization.

In general, any substituent that has a lone pair of electrons on the atom bound to the aromatic ring has an electron-donating resonance effect and is thus an ortho–para director. Halogens, hydroxyl groups, and amino groups are therefore all ortho–para directors:

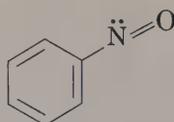


Whether a given substituent is activating or deactivating depends on the relative strengths of its inductive and resonance effects. For example, both halogens (*o*- and *p*-directing deactivators) and hydroxyl (*o*- and *p*-directing activators) show the same two effects—inductive electron withdrawal and resonance electron donation—but the relative strengths of the effects differ in each case. Halogens have a strong electron-withdrawing inductive effect but a weak electron-donating resonance effect and are thus deactivators. Hydroxyl and amino groups have weak electron-withdrawing inductive effects but have strong electron-donating resonance effects and are thus activators.



PROBLEM .....

- 16.20 The nitroso group,  $-\text{N}=\text{O}$ , is one of the very few nonhalogens that is an ortho- and para-directing deactivating group. Draw resonance structures of intermediates in ortho and para electrophilic attack on nitrosobenzene, and explain why they are favored over the intermediate from meta attack.



Nitrosobenzene

.....

## 16.7 Substituent Effects in Aromatic Substitution: A Summary

Orientation and reactivity in electrophilic aromatic substitution are controlled by the interplay of two factors: resonance effects and inductive effects. Different substituents behave differently, depending on the direction and strength of the two effects (Table 16.2). We can summarize the results as follows:

1. *Alkyl groups* Moderate electron-donating inductive effect; no resonance effect. The net result is that alkyl groups are activating and ortho- and para-directing.
2. *Hydroxyl and amino groups (and derivatives)* Strong electron-donating resonance effect; moderate electron-withdrawing inductive effect. The net result is that these groups are activating and ortho- and para-directing.
3. *Halogens* Strong electron-withdrawing inductive effect; moderate electron-donating resonance effect. The net result is that halogens deactivate the ring but are ortho- and para-directing.
4. *Nitro, cyano, carbonyl, and similar groups* Strong electron-withdrawing resonance effect; strong electron-withdrawing inductive effect. The net result is that these groups are meta-directing and deactivating.

TABLE 16.2 Substituent Effects in Electrophilic Aromatic Substitution

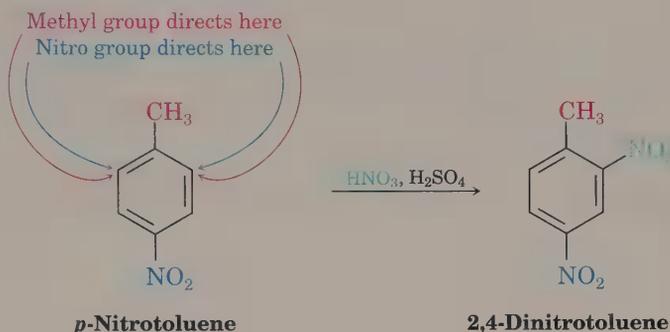
Substituent	Reactivity	Orientation	Inductive effect	Resonance effect
—CH <sub>3</sub>	Activating	Ortho, para	Weak; electron-donating	None
— $\ddot{\text{O}}\text{H}$ — $\ddot{\text{N}}\text{H}_2$	Activating	Ortho, para	Weak; electron-withdrawing	Strong; electron-donating
— $\ddot{\text{F}}:$ , — $\ddot{\text{Cl}}:$ — $\ddot{\text{Br}}:$ , — $\ddot{\text{I}}:$	Deactivating	Ortho, para	Strong; electron-withdrawing	Weak; electron-donating
— $\overset{+}{\text{N}}(\text{CH}_3)_3$	Deactivating	Meta	Strong; electron-withdrawing	None
—NO <sub>2</sub> , —CN —CHO, —CO <sub>2</sub> CH <sub>3</sub> —COCH <sub>3</sub>	Deactivating	Meta	Strong; electron-withdrawing	Strong; electron-withdrawing

## 16.8 Trisubstituted Benzenes: Additivity of Effects

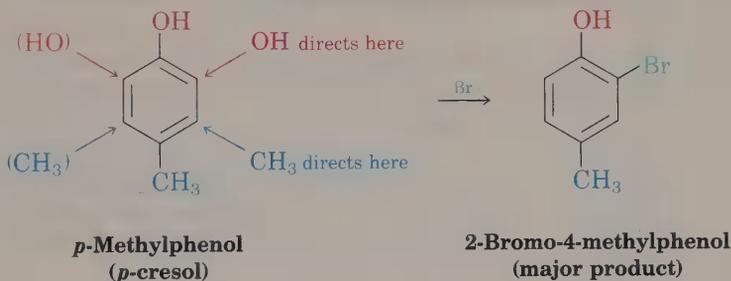
Further electrophilic substitution of a disubstituted benzene is governed by the same resonance and inductive effects just discussed. The only difference is that it's now necessary to consider the additive effects of two different

groups. In practice, this isn't as difficult as it sounds; three rules are usually sufficient:

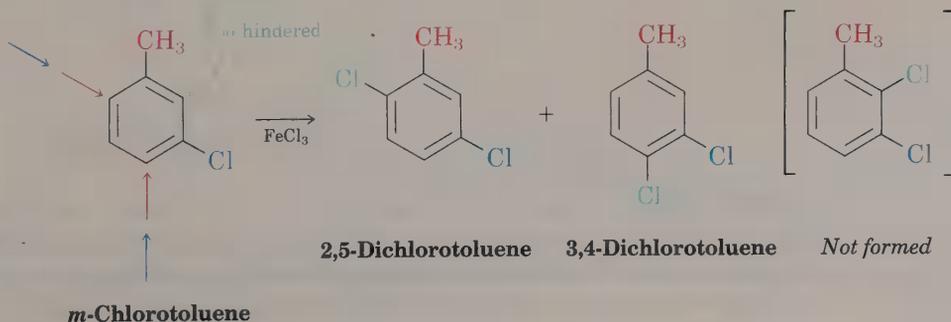
1. If the directing effects of the two groups reinforce each other, there's no problem. In *p*-nitrotoluene, for example, both the methyl and the nitro group direct further substitution to the same position (ortho to the methyl = meta to the nitro). A single product is thus formed by electrophilic substitution.



2. If the directing effects of the two groups oppose each other, the more powerful activating group has the dominant influence, but mixtures of products often result. For example, bromination of *p*-methylphenol yields largely 2-bromo-4-methylphenol, since hydroxyl is a more powerful activator than methyl is.



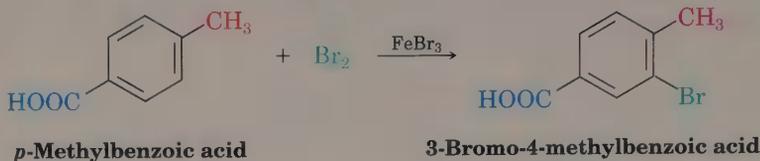
3. Further substitution rarely occurs between the two groups in a meta-disubstituted compound because this site is too hindered for reaction to occur easily.



## PRACTICE PROBLEM.....

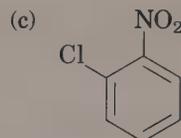
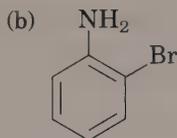
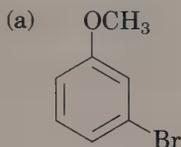
What product would you expect from bromination of *p*-methylbenzoic acid?

**Solution** The carboxyl group (–COOH) is a meta director, and the methyl group is an ortho and para director. Both groups therefore direct bromination to the position next to the methyl group, yielding 3-bromo-4-methylbenzoic acid.



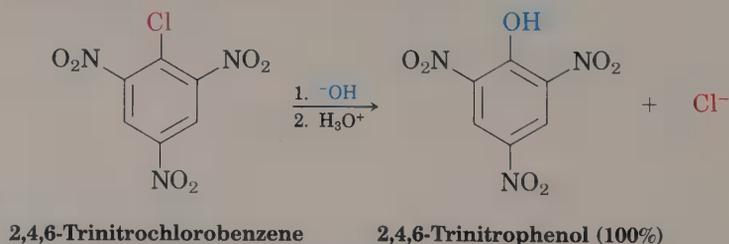
## PROBLEM.....

16.21 Where would you expect electrophilic substitution to occur in these substances?



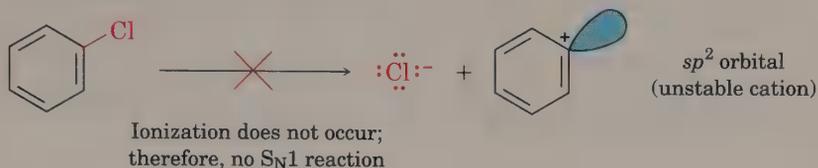
## 16.9 Nucleophilic Aromatic Substitution

The electrophilic substitutions just discussed are the most important and useful reactions of the aromatic ring. In certain cases, however, aromatic substitution can also occur by a *nucleophilic* mechanism. Nucleophilic substitutions aren't characteristic of all aromatic rings, though. Only aryl halides that have electron-withdrawing substituents undergo **nucleophilic aromatic substitution**. For example, 2,4,6-trinitrochlorobenzene reacts with aqueous sodium hydroxide at room temperature to give 2,4,6-trinitrophenol in 100% yield. The nucleophile, hydroxide ion, has substituted for chloride ion:

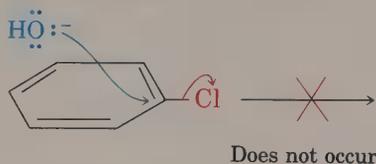


How does this reaction take place? Although it appears similar to the S<sub>N</sub>1 and S<sub>N</sub>2 nucleophilic substitution reactions of alkyl halides (Chapter 11), nucleophilic aromatic substitution is in fact quite different, because aryl halides are inert to substitution by S<sub>N</sub>1 and S<sub>N</sub>2 mechanisms.

Aryl halides don't undergo  $S_N1$  reactions because aryl cations are relatively unstable. The dissociation of an aryl halide is therefore energetically unfavorable and does not occur easily.



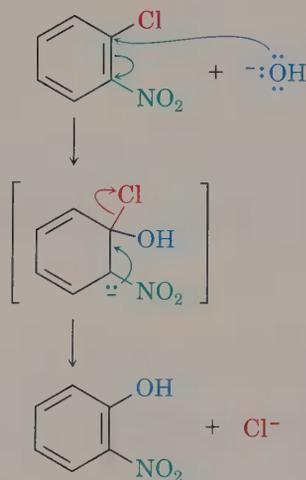
Aryl halides don't undergo  $S_N2$  reactions because the halo-substituted carbon atom is sterically shielded from back-side attack by the aromatic ring. In order for a nucleophile to attack an aryl halide, it would have to approach directly through the aromatic ring and invert the stereochemistry of the aromatic ring—a geometric impossibility.



Nucleophilic aromatic substitutions proceed by the *addition-elimination* mechanism shown in Figure 16.20. The attacking nucleophile first adds to the electron-deficient aryl halide, forming a negatively charged intermediate called a Meisenheimer<sup>3</sup> complex, and halide ion is then eliminated in the second step.

Nucleophilic addition of hydroxide ion to the electron-poor aromatic ring takes place yielding a stabilized carbanion intermediate.

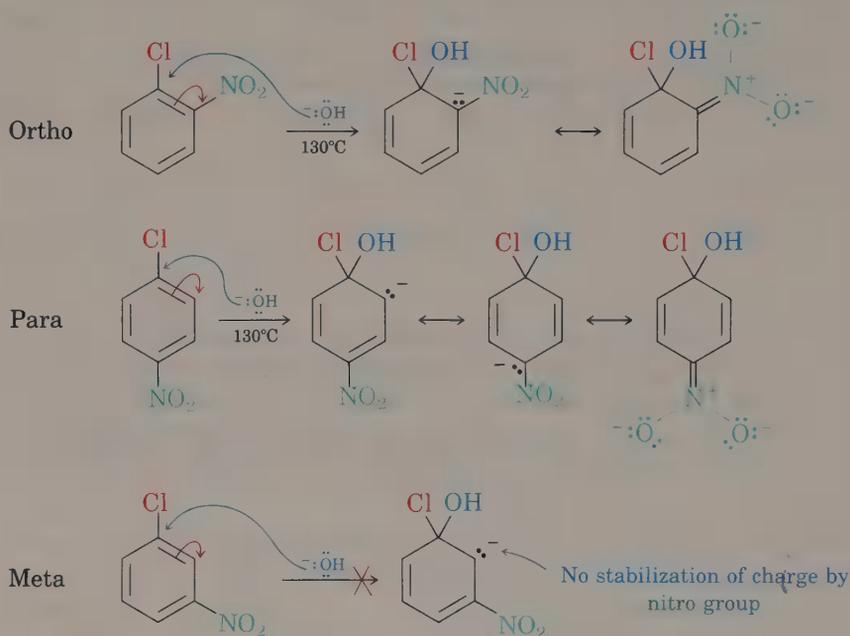
The carbanion intermediate undergoes elimination of chloride ion in a second step to give the substitution product.



**Figure 16.20** Mechanism of aromatic nucleophilic substitution. The reaction occurs in two steps and involves a stabilized carbanion intermediate.

<sup>3</sup>Jacob Meisenheimer (1876–1934); b. Greisheim; Ph.D. Munich; professor, universities of Berlin, Griefswald, and Tübingen.

Nucleophilic aromatic substitution occurs only if the halobenzene has electron-withdrawing substituents in the ortho and/or para positions. The more substituents there are, the faster the reaction goes. As shown in Figure 16.21, only ortho and para electron-withdrawing substituents can stabilize the anion intermediate through resonance; a meta substituent offers no such resonance stabilization. Thus, *p*-chloronitrobenzene and *o*-chloronitrobenzene react with hydroxide ion at 130°C to yield substitution products, but *m*-chloronitrobenzene is inert to hydroxide ion.

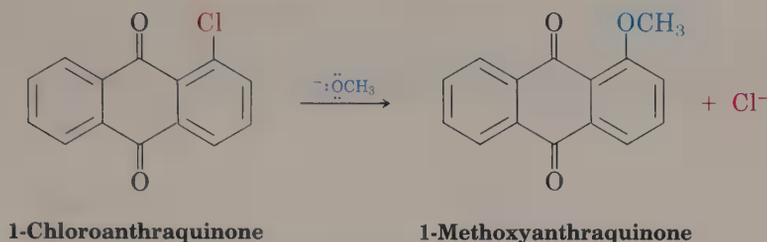


**Figure 16.21** Nucleophilic aromatic substitution on nitrochlorobenzenes. Only the ortho and para isomers undergo reaction.

Note the different characteristics of aromatic substitution through electrophilic and nucleophilic pathways: The electron-withdrawing groups that *deactivate* rings for electrophilic substitution (nitro, carbonyl, cyano, and so on) *activate* them for nucleophilic substitution. What's more, these groups are meta directors in electrophilic substitution, but ortho–para directors in nucleophilic substitution.

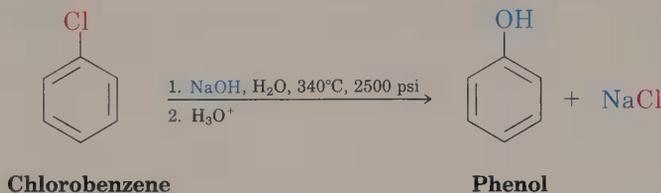
**PROBLEM** .....

- 16.22** Propose a mechanism to account for the observation that 1-chloroanthraquinone reacts with methoxide ion to give the substitution product 1-methoxyanthraquinone.

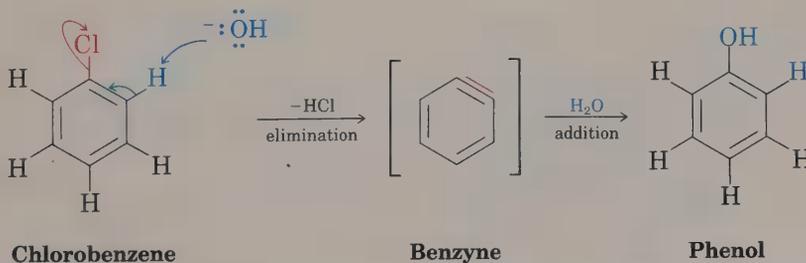


## 16.10 Benzyne

Halobenzenes without electron-withdrawing substituents are inert to nucleophiles under normal conditions. At very high temperature and pressure, however, even chlorobenzene can be forced to react. Chemists at the Dow Chemical Company discovered in 1928 that phenol could be prepared on a large industrial scale by treatment of chlorobenzene with dilute aqueous sodium hydroxide at 340°C under 2500 psi (pounds per square inch) pressure.

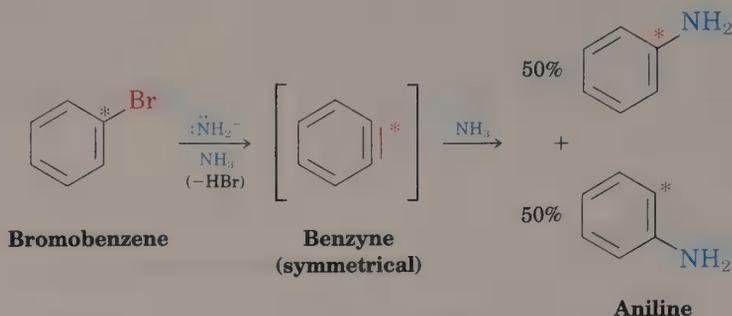


This phenol synthesis is different from the other nucleophilic aromatic substitution reactions just studied in that it takes place by an *elimination–addition* mechanism rather than an addition–elimination. Strong base first causes the *elimination* of HX from halobenzene, yielding a highly reactive **benzyne** intermediate, and a nucleophile then *adds* to benzyne in a second step to yield the product. The two steps are similar to those in other nucleophilic aromatic substitutions, but their order is reversed: addition before elimination for the usual reaction versus elimination before addition for the benzyne reaction.

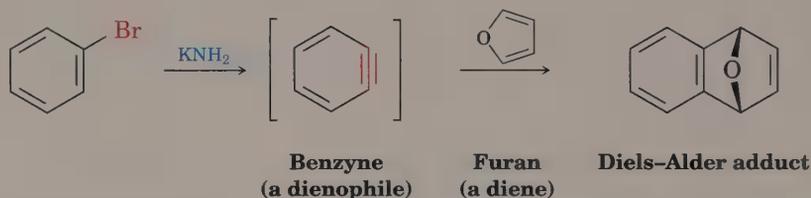


Evidence in support of the benzyne mechanism has been obtained by studying the reaction between bromobenzene and the strong base potassium amide. When bromobenzene labeled with a radioactive  $^{14}C$  carbon atom at the 1 position is used, the substitution product has the label scrambled

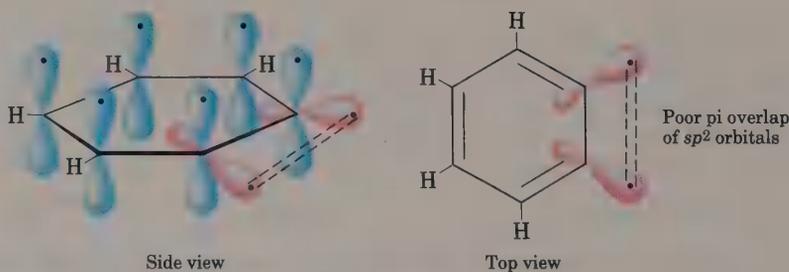
between positions 1 and 2. This labeling result requires that the reaction proceed through a symmetrical intermediate in which positions 1 and 2 are equivalent—a requirement that only benzyne can meet:



Further evidence for a benzyne intermediate comes from trapping experiments. Although benzyne is much too reactive to be isolated as a pure compound, it can be intercepted in a Diels–Alder reaction if a diene such as furan is present when benzyne is generated. This is just the kind of behavior we would expect for so strained and reactive a species.



The electronic structure of benzyne shown in Figure 16.22 can be compared to that of a highly distorted alkyne. Although a normal alkyne triple bond consists of a sigma bond formed by  $sp-sp$  overlap and two mutually perpendicular pi bonds formed by  $p-p$  overlap, the benzyne triple bond consists of a sigma bond formed by  $sp^2-sp^2$  overlap, one pi bond formed by  $p-p$  overlap, and one pi bond formed by  $sp^2-sp^2$  overlap. The latter pi bond is in the plane of the ring and is very weak because of poor orbital overlap.



**Figure 16.22** An orbital picture of benzyne. The benzyne carbons are  $sp^2$  hybridized, and the “third” bond results from weak overlap of two adjacent  $sp^2$  orbitals.

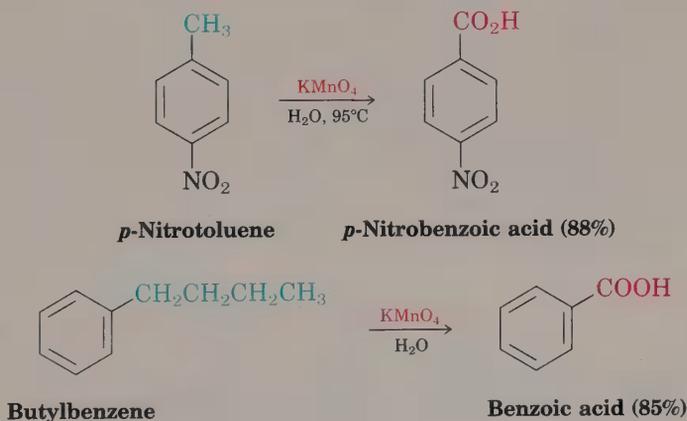
## PROBLEM .....

- 16.23 Account for the fact that treatment of *p*-bromotoluene with NaOH at 300°C yields a mixture of two products, but treatment of *m*-bromotoluene with NaOH yields a mixture of three products.
- .....

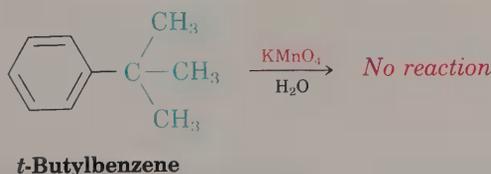
## 16.11 Oxidation of Aromatic Compounds

### OXIDATION OF ALKYL BENZENE SIDE CHAINS

Despite its unsaturation, the benzene ring is normally inert to strong oxidizing agents such as potassium permanganate and sodium dichromate, reagents that will cleave alkene carbon-carbon bonds (Section 7.8). It turns out, however, that the presence of the aromatic ring has a dramatic effect on alkyl-group side chains. Alkyl side chains are readily attacked by oxidizing agents and are converted into carboxyl groups,  $-\text{COOH}$ . The net effect of side-chain oxidation is the conversion of an alkylbenzene into a benzoic acid,  $\text{Ar-R} \rightarrow \text{Ar-COOH}$ . For example, both *p*-nitrotoluene and butylbenzene are oxidized by aqueous  $\text{KMnO}_4$  in high yield to give the corresponding benzoic acids.

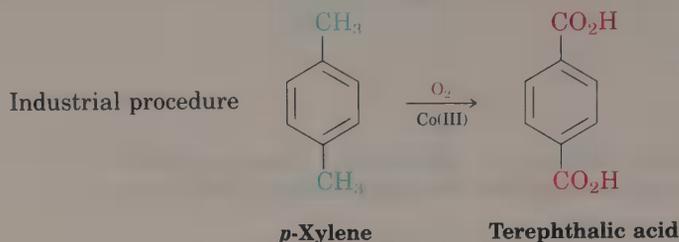


The exact mechanism of the reaction is not fully understood but probably involves attack on side-chain C-H bonds at the position next to the aromatic ring to form intermediate benzylic radicals. *tert*-Butylbenzene has no benzylic hydrogens, however, and is therefore inert.



A similar oxidation is employed industrially for the preparation of terephthalic acid, used in production of polyester fibers. Approximately five

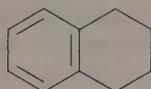
million tons per year of *p*-xylene are oxidized in this manner, using air as the oxidant and Co(III) salts as catalyst.



PROBLEM .....

16.24 What aromatic products would you expect to obtain from the  $\text{KMnO}_4$  oxidation of these substances?

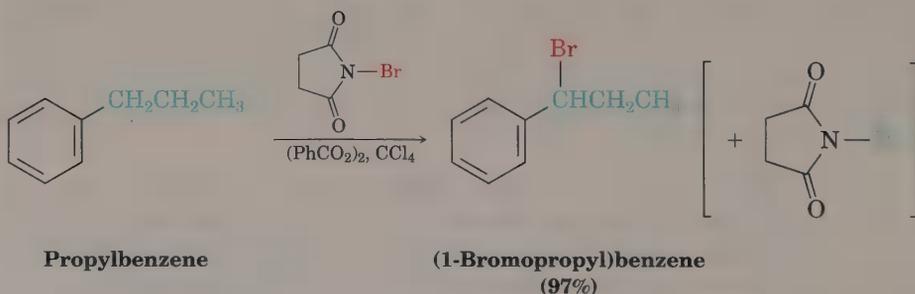
(a) Tetralin,



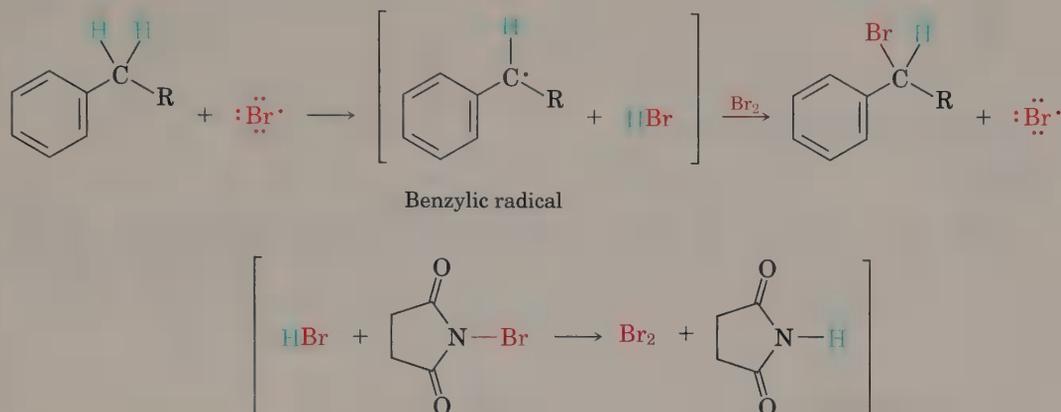
(b) *m*-Nitroisopropylbenzene

BROMINATION OF ALKYL BENZENE SIDE CHAINS

Another kind of side-chain oxidation takes place when alkylbenzenes are treated with *N*-bromosuccinimide (NBS). *N*-Bromosuccinimide reacts with alkylbenzenes to brominate the benzylic position through a radical chain mechanism. For example, propylbenzene gives (1-bromopropyl)benzene in 97% yield on reaction with NBS in the presence of benzoyl peroxide,  $(\text{PhCO}_2)_2$ , as a radical initiator. Note that bromination occurs exclusively in the benzylic position and does not give a mixture of products.

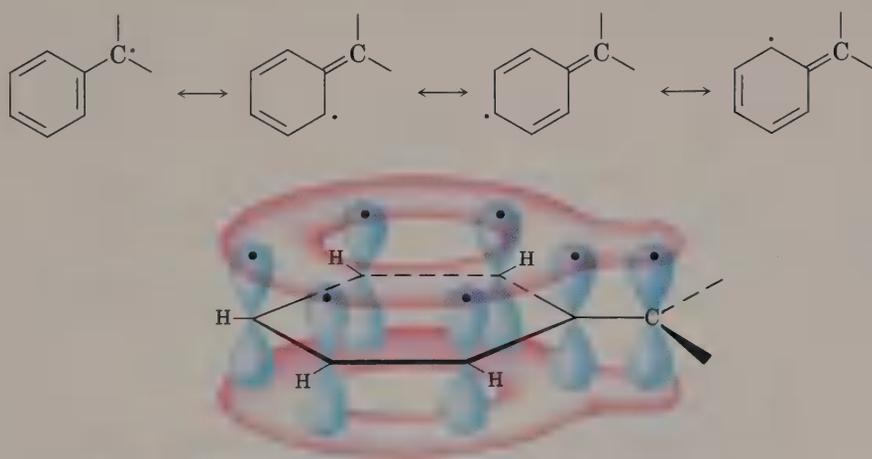


The mechanism of benzylic bromination is similar to that discussed previously for allylic bromination of alkenes (Section 10.5) and involves abstraction of a benzylic hydrogen atom of the alkylbenzene to generate an intermediate benzyl radical. The stabilized radical then reacts with  $\text{Br}_2$  to yield product and a bromine radical, which cycles back into the reaction to carry on the chain. The  $\text{Br}_2$  necessary for reaction with the benzyl radical is produced by a concurrent reaction of HBr with NBS, as shown in Figure 16.23.



**Figure 16.23** Mechanism of benzylic bromination with *N*-bromosuccinimide.

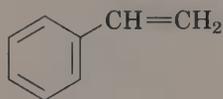
Reaction occurs exclusively at the benzylic position because the benzylic radical is highly stabilized by resonance. Figure 16.24 shows an orbital view of the benzyl radical, indicating how the radical is stabilized by overlap of its *p* orbital with the ring pi electron system.



**Figure 16.24** An orbital picture of a benzylic radical, showing the overlap of a side-chain orbital with the aromatic ring orbitals.

**PROBLEM** .....

- 16.25** Styrene, the simplest alkenylbenzene, is prepared commercially for use in plastics manufacture by dehydrogenation of ethylbenzene over a special catalyst. How might you prepare styrene from benzene using reactions you have studied?



Styrene

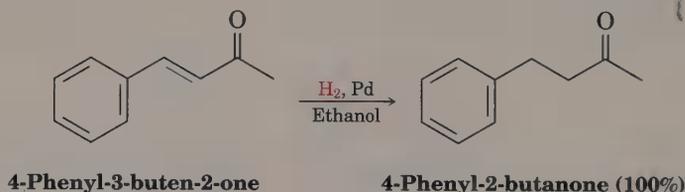
## PROBLEM .....

- 16.26** Refer to Table 5.4 for a quantitative idea of the stability of a benzyl radical. Approximately how much more stable (in kcal/mol) is the benzyl radical than a primary alkyl radical? How does a benzyl radical compare in stability to an allyl radical?
- .....

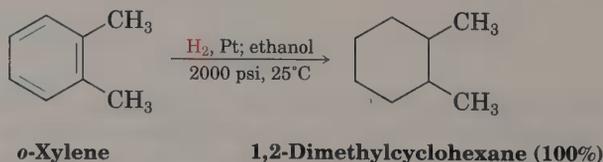
## 16.12 Reduction of Aromatic Compounds

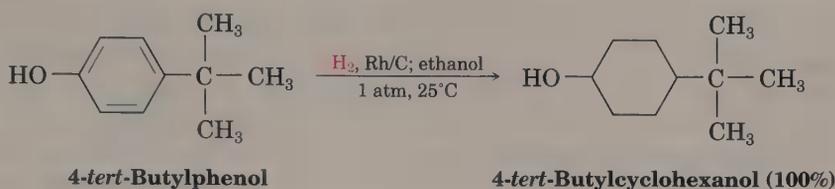
### CATALYTIC HYDROGENATION OF AROMATIC RINGS

Just as aromatic rings are inert to oxidation under normal conditions, they're also inert to catalytic hydrogenation under conditions that reduce ordinary alkene double bonds. The usual platinum and palladium alkene hydrogenation catalysts don't affect aromatic rings under most conditions, and it's therefore possible to selectively reduce alkene double bonds in the presence of aromatic rings. For example, 4-phenyl-3-buten-2-one is selectively reduced to 4-phenyl-2-butanone when the reaction is carried out at room temperature and atmospheric pressure using a palladium catalyst. Neither the benzene ring nor the ketone carbonyl group is affected.



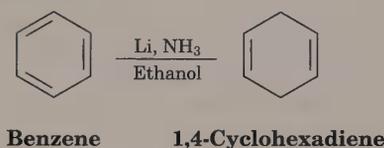
In order to hydrogenate an aromatic ring, it's necessary either to use an ordinary platinum catalyst with hydrogen gas at several hundred atmospheres pressure or to use a more powerful catalyst such as rhodium on carbon. Under these conditions, aromatic rings are readily reduced to cyclohexanes. For example, *o*-xylene yields 1,2-dimethylcyclohexane, and 4-*tert*-butylphenol gives 4-*tert*-butylcyclohexanol in 100% yield.





## BIRCH REDUCTION OF AROMATIC RINGS

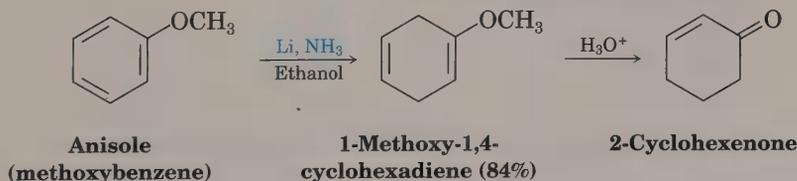
Although aromatic rings are less reactive than alkenes toward catalytic hydrogenation, they are more reactive toward reduction by alkali metals. Treatment of benzene with lithium or sodium metal in a mixed liquid ammonia/ethanol solvent leads to rapid reduction of the aromatic ring and production of 1,4-cyclohexadiene.



Known as a **Birch reduction**, the reaction of aromatic compounds with alkali metals is mechanistically similar to the reduction of alkynes discussed in Section 8.7. As with alkyne reduction, the process is initiated by donation of an electron from the alkali metal to the pi electron system of the aromatic ring. Having gained an electron, the product of this first step is both an anion and a radical, and is stabilized by resonance among several Kekulé structures.

The anion radical intermediate is immediately protonated by the ethanol solvent to give a radical, which is further reduced by a second equivalent of lithium to yield a carbanion. Transfer of a second proton from ethanol to the carbanion then gives the cyclohexadiene product (Figure 16.25).

Alkyl-substituted benzenes sometimes give mixtures of isomeric cyclohexadiene products on Birch reduction, but methoxybenzenes reduce cleanly to a single isomer. The resultant methoxy-substituted product is useful because it can be converted into a cyclohexenone by acid-catalyzed hydrolysis.

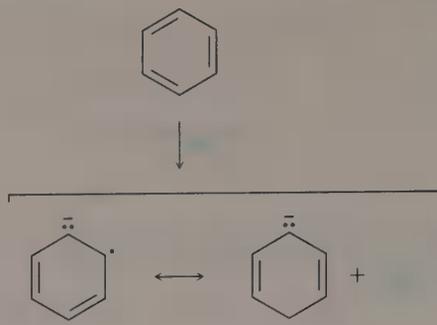


### PROBLEM.....

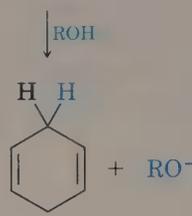
16.27 What two products would you expect from Birch reduction of *m*-xylene?

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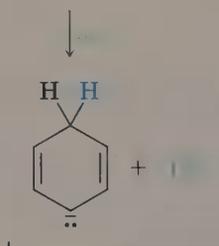
Lithium metal donates an electron to the pi system of the aromatic ring, giving an anion radical.



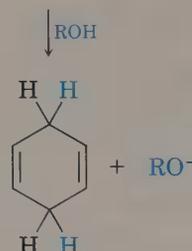
The anion radical is protonated by ethanol, giving a radical.



The radical is reduced by a second equivalent of lithium to generate a carbanion.



The carbanion is protonated by ethanol to produce the neutral reduction product.

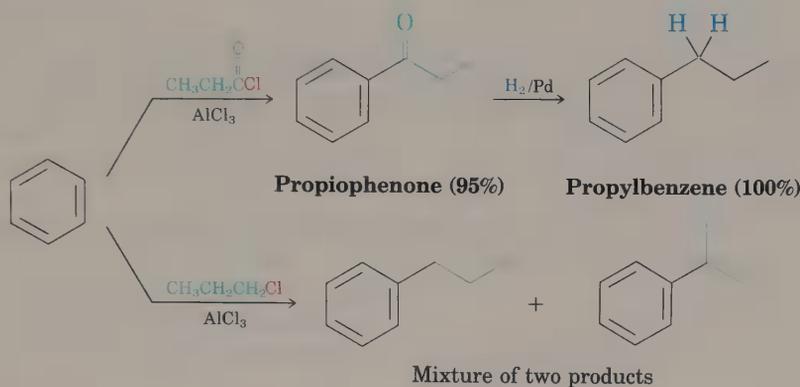


**Figure 16.25** Mechanism of the Birch reduction of benzene with lithium metal in ammonia/ethanol.

## REDUCTION OF ARYL ALKYL KETONES

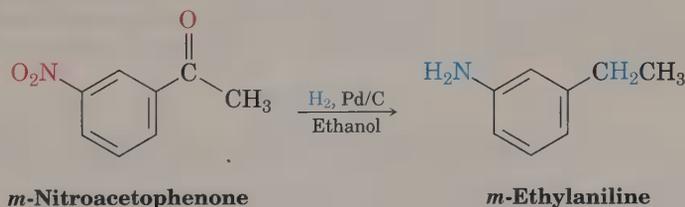
Just as the presence of an aromatic ring activates a neighboring benzylic C–H position toward oxidation, it also activates a neighboring carbonyl group toward reduction. Thus, an aryl alkyl ketone prepared by Friedel–Crafts acylation of an aromatic ring can be converted into an alkylbenzene by catalytic hydrogenation over a palladium catalyst. For example, propio-

phenone is reduced to propylbenzene in 100% yield by catalytic hydrogenation. Since the net effect of Friedel–Crafts acylation followed by reduction is the preparation of a primary alkylbenzene, this two-step sequence of reactions makes it possible to circumvent the carbocation rearrangement problems associated with direct Friedel–Crafts alkylation using primary alkyl halides (Figure 16.26).



**Figure 16.26** Use of the Friedel–Crafts acylation reaction to prepare a straight-chain alkylbenzene.

Note that the conversion of a carbonyl group into a methylene ( $\text{C}=\text{O} \rightarrow \text{CH}_2$ ) by catalytic hydrogenation is limited to *aryl* alkyl ketones. The presence of the neighboring aromatic ring is necessary to increase the reactivity of the carbonyl group toward hydrogenation; dialkyl ketones are not reduced under these conditions. It should also be pointed out that the catalytic reduction of aryl alkyl ketones is not compatible with the presence of a nitro substituent on the aromatic ring, since nitro groups are reduced to amino groups under the reaction conditions. We'll see a more general method for reducing ketone carbonyl groups to yield alkanes in Section 19.13.



**PROBLEM** .....

**16.28** Show how you would prepare diphenylmethane,  $(\text{Ph})_2\text{CH}_2$ , from benzene and an appropriate acid chloride.

.....

## 16.13 Synthesis of Substituted Benzenes

One of the surest ways to acquire a command of organic chemistry is to work synthesis problems. The ability to plan a successful multistep synthesis of a complex molecule requires a working knowledge of the uses and limitations of many hundreds of organic reactions. Not only must you know *which* reactions to use, you must also know *when* to use them. The order in which reactions are carried out is often critical to the success of the overall scheme.

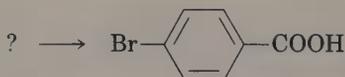
The ability to plan a sequence of reactions in the right order is particularly valuable in the synthesis of substituted aromatic rings, where the introduction of one substituent is strongly affected by the directing effects of other substituents. Planning syntheses of substituted aromatic compounds is therefore an excellent way to gain facility with the many reactions learned in the past two chapters.

During our earlier discussion of the strategies that can be used in working synthesis problems (Section 8.11), we said that it's usually best to work problems backward. Look at the target molecule and ask yourself, "What is an immediate precursor of this compound?" Choose a likely answer and continue working backward, one step at a time, until you arrive at a simple starting material. Let's try some examples.

### PRACTICE PROBLEM.....

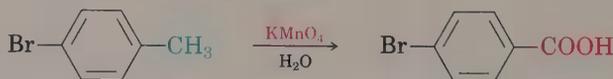
Synthesize *p*-bromobenzoic acid from benzene.

**Solution** Ask yourself, "What is an immediate precursor of *p*-bromobenzoic acid?"



*p*-Bromobenzoic acid

There are only two substituents on the ring, a carboxyl group (COOH), which is meta-directing, and a bromine, which is ortho-para-directing. We can't brominate benzoic acid, because the wrong isomer (*m*-bromobenzoic acid) would be produced. We know, however, that oxidation of alkylbenzene side chains yields benzoic acids. Thus, an immediate precursor of our target molecule might be *p*-bromotoluene.

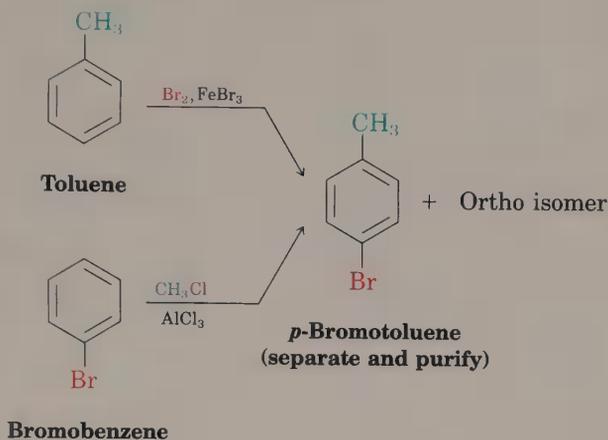


*p*-Bromotoluene

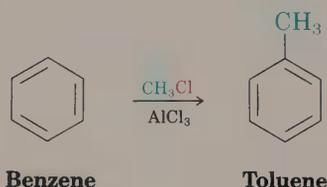
*p*-Bromobenzoic acid

Next ask yourself, "What is an immediate precursor of *p*-bromotoluene?" Perhaps toluene is an immediate precursor, since the methyl group would direct bromination to the ortho and para positions, and the isomeric products could be separated. Alternatively, bromobenzene might be an immediate precursor, since we could carry out a Friedel-Crafts methylation and obtain para product. *Both answers are satisfactory*, although, in view of the difficulties often observed with polyalkyl-

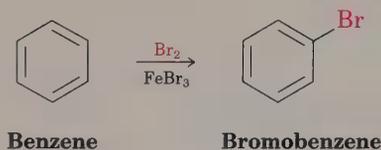
ation in Friedel–Crafts reactions, bromination of toluene may well be the more efficient route.



“What is an immediate precursor of toluene?” Benzene, which could be methylated in a Friedel–Crafts reaction.



Alternatively, “What is an immediate precursor of bromobenzene?” Benzene, which could be brominated.

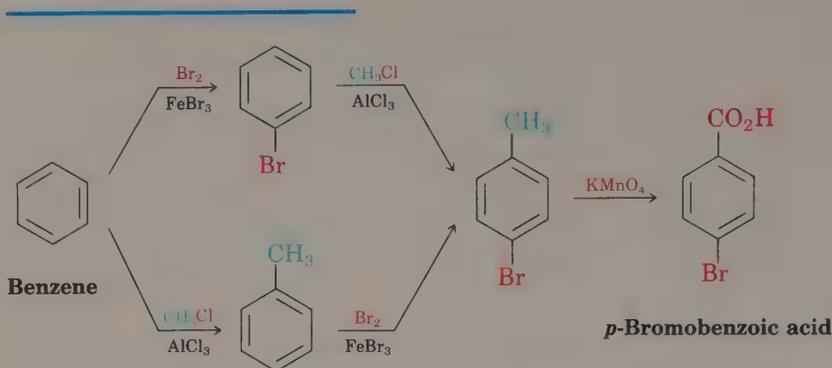


This backward synthetic (*retrosynthetic*) analysis has provided two valid routes from benzene to *p*-bromobenzoic acid (Figure 16.27).

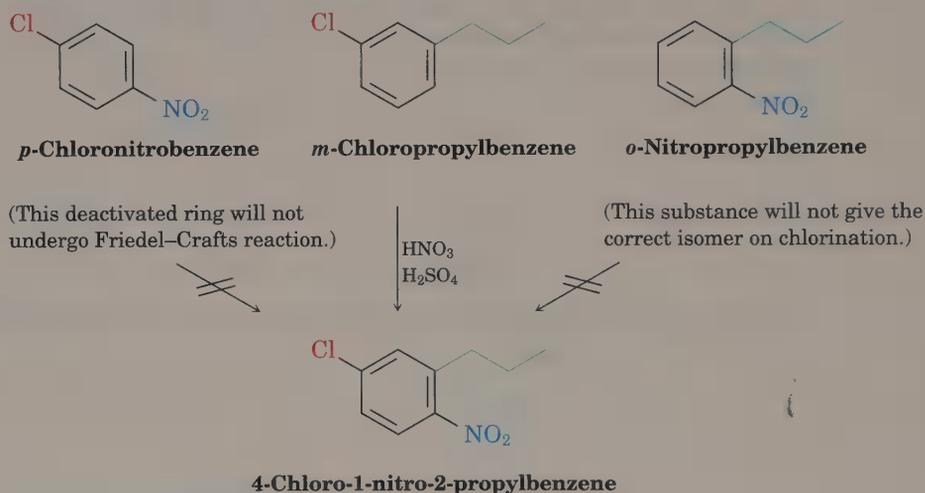
#### PRACTICE PROBLEM.....

Propose a synthesis of 4-chloro-1-nitro-2-propylbenzene from benzene.

**Solution** “What is an immediate precursor of the target?” Since the final step will involve introduction of one of three groups—chloro, nitro, or propyl—we have to consider three possibilities. Of the three, we know that chlorination of *o*-nitropropylbenzene can’t be used because the reaction would occur at the wrong position. Similarly, a Friedel–Crafts reaction can’t be used as the final step because these reactions don’t work on nitro-substituted (deactivated) benzenes. Thus, the immediate precursor of our desired product is probably *m*-chloropropylbenzene, which can be nitrated. This nitration gives a mixture of product isomers, which must then be separated (Figure 16.28).

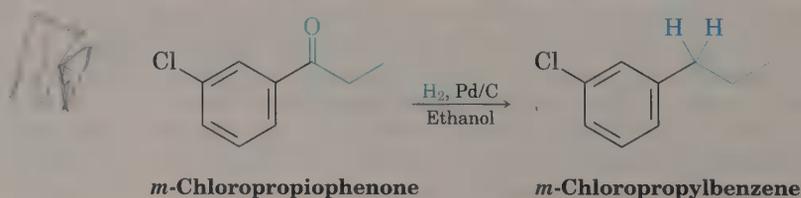


**Figure 16.27** Two routes for the synthesis of *p*-bromobenzoic acid from benzene.



**Figure 16.28** Possible routes for the synthesis of 4-chloro-1-nitro-2-propylbenzene.

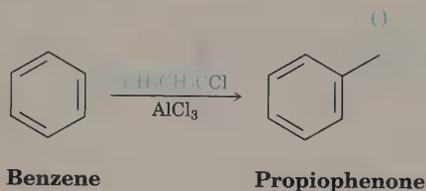
“What is an immediate precursor of *m*-chloropropylbenzene?” Since the two substituents have a meta relationship, the *first* substituent placed on the ring must be a meta director so that the second substitution will take place at the proper position. Furthermore, since primary alkyl groups such as propyl can’t be introduced directly by Friedel-Crafts alkylation, the precursor of *m*-chloropropylbenzene is probably *m*-chloropropiophenone, which could undergo catalytic hydrogenation of the acyl group.



“What is an immediate precursor of *m*-chloropropiophenone?” Perhaps propiophenone, which could be chlorinated.



“What is an immediate precursor of propiophenone?” Benzene, which could undergo Friedel–Crafts acylation with propanoyl chloride and  $\text{AlCl}_3$ .



The final synthesis is a four-step route from benzene:



Planning organic syntheses has been compared to playing chess. There are no tricks; all that's required is a knowledge of the allowable moves (the organic reactions) and the discipline to work backward and to evaluate carefully the consequences of each move. Practicing is not always easy, but there's no surer way to learn organic chemistry.

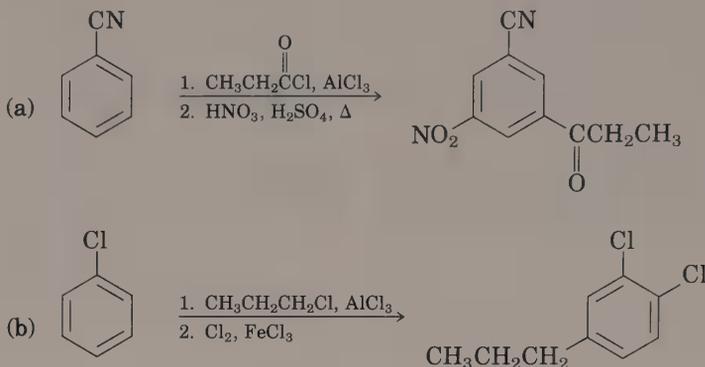
PROBLEM.....

16.29 Propose syntheses of these substances from benzene:

- (a) *m*-Chloronitrobenzene
- (b) *m*-Chloroethylbenzene
- (c) *p*-Chloropropylbenzene

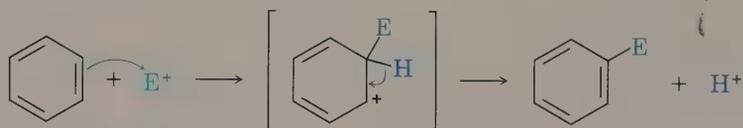
## PROBLEM .....

**16.30** In planning syntheses, it is as important to know what not to do as to know what to do. As written, the following reaction schemes have flaws that make their success unlikely. What is wrong with each one?



## 16.14 Summary and Key Words

**Electrophilic aromatic substitution** is the most important reaction of aromatic compounds. The reaction takes place in two steps—initial reaction of an electrophile,  $E^+$ , with the aromatic ring, followed by loss of a proton from the resonance-stabilized carbocation intermediate to regenerate the aromatic ring:



Many different substituents can be introduced onto the ring by this process. **Bromination, chlorination, iodination, nitration, sulfonation, alkylation,** and **acylation** can all be carried out with the proper choice of reagent.

**Friedel-Crafts alkylation** and **acylation**, which involve reaction of an aromatic with carbocation electrophiles, are particularly useful but are limited in several ways:

1. Only alkyl halides and acyl halides can be used. Vinylic and aryl halides don't react.
2. The aromatic ring must be at least as reactive as a halobenzene. Strongly deactivated rings don't react.
3. Polyalkylation often occurs in Friedel-Crafts alkylation, since the product alkylbenzene is more reactive than the starting material.
4. Carbocation rearrangements can occur during Friedel-Crafts alkylation, particularly when primary alkyl halides are used.

Substituents on the benzene ring affect both the reactivity of the ring toward further substitution and the orientation of further substitution. Groups can be classified into three categories: **ortho- and para-directing activators**, **ortho- and para-directing deactivators**, and **meta-directing deactivators**.

Substituent effects are due to an interplay of resonance and inductive effects. **Resonance effects** are transmitted through  $p$  orbitals when the atom directly attached to the aromatic ring has a  $p$  orbital that can overlap the aromatic ring pi system; **inductive effects** are transmitted through sigma bonds.

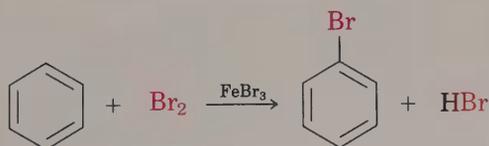
When electrophilic substitution is carried out on a disubstituted benzene, both of the groups already present exert their orienting effects independently. If both groups direct substitution toward the same position, reaction occurs at that site. If the groups have conflicting directional effects, the more powerful activating substituent controls the reaction.

In special cases, halobenzenes undergo **nucleophilic aromatic substitution** through either of two mechanisms. If the halobenzene has strongly electron-withdrawing substituents in the ortho and/or para position, substitution occurs by addition of a nucleophile to the ring, followed by elimination of halide from the intermediate anion. If the halobenzene is not activated by electron-withdrawing substituents, nucleophilic substitution can occur by elimination of HX, followed by addition of a nucleophile to the intermediate **benzyne**.

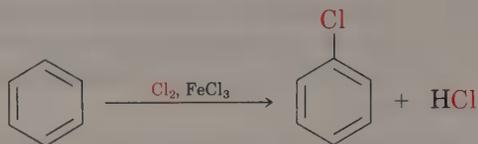
The side chain of alkylbenzenes has unique reactivity because of the neighboring aromatic ring. Thus, the benzylic position can be brominated by reaction with *N*-bromosuccinimide, and the entire side chain can be degraded to a carboxylic acid by oxidation with aqueous potassium permanganate. Although aromatic rings are much less reactive than isolated alkene double bonds, they can nevertheless be reduced to cyclohexanes by hydrogenation over a platinum or rhodium catalyst and to 1,4-cyclohexadienes by **Birch reduction** with an alkali metal. The neighboring aromatic ring also allows aryl alkyl ketones to be reduced to alkylbenzenes by hydrogenation over a platinum catalyst. The net effect of Friedel–Crafts acylation followed by catalytic hydrogenation is the synthesis of a straight-chain alkylbenzene.

## 16.15 Summary of Reactions

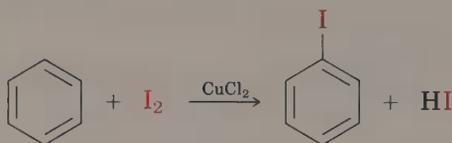
1. Electrophilic aromatic substitution
  - a. Bromination (Section 16.1)



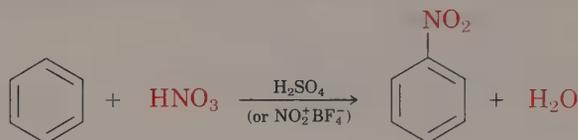
## b. Chlorination (Section 16.2)



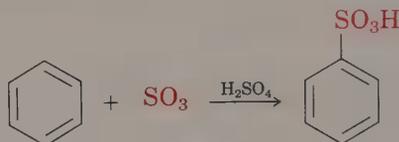
## c. Iodination (Section 16.2)



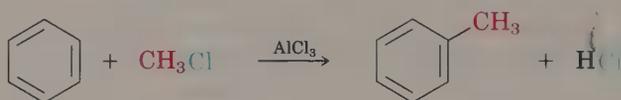
## d. Nitration (Section 16.2)



## e. Sulfonation (Section 16.2)



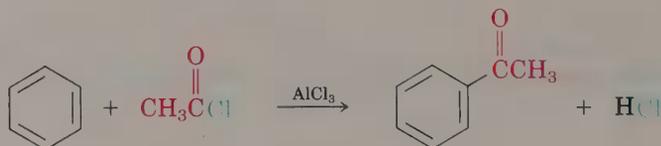
## f. Friedel–Crafts alkylation (Section 16.3)



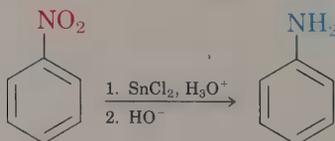
**Aromatic ring:** Must be at least as reactive as a halobenzene. Deactivated rings do not react.

**Alkyl halide:** Can be methyl, ethyl, 2°, or 3°; primary halides undergo carbocation rearrangement.

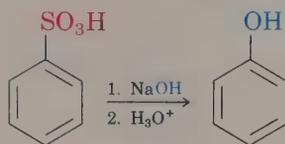
## g. Friedel–Crafts acylation (Section 16.4)



## 2. Reduction of aromatic nitro groups (Section 16.2)

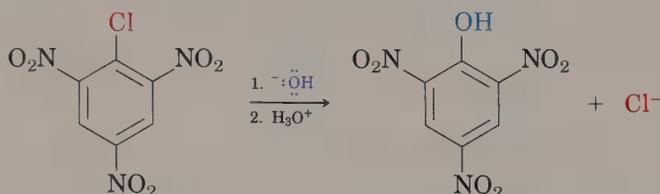


## 3. Alkali fusion of aromatic sulfonates (Section 16.2)

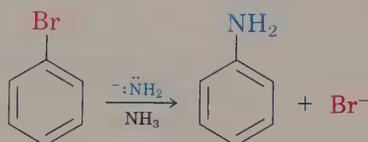


## 4. Nucleophilic aromatic substitution

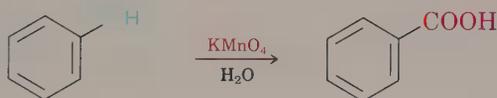
## a. By addition/elimination to activated aryl halides (Section 16.9)



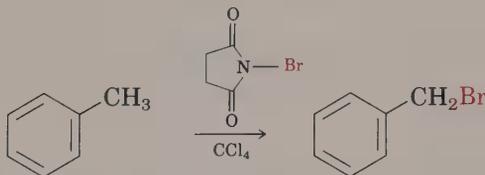
## b. By benzyne intermediate for unactivated aryl halides (Section 16.10)



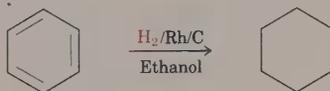
## 5. Oxidation of alkylbenzene side chains (Section 16.11)



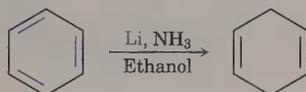
Reaction occurs with 1° and 2°, but not 3°, alkyl side chains.

6. *N*-Bromosuccinimide bromination of alkylbenzenes (Section 16.11)

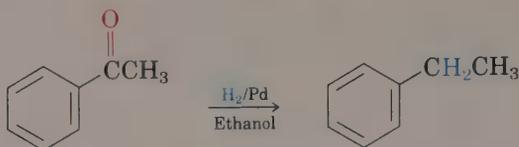
## 7. Catalytic hydrogenation of aromatic rings (Section 16.12)



## 8. Birch reduction of aromatic rings (Section 16.12)



## 9. Reduction of aryl alkyl ketones (Section 16.12)



Reaction is specific for aryl alkyl ketones; dialkyl ketones are not affected.

## ADDITIONAL PROBLEMS .....

- 16.31** Predict the major product(s) of mononitration of these substances. Which react faster than benzene, and which slower?
- (a) Bromobenzene                      (b) Benzonitrile                      (c) Benzoic acid  
(d) Nitrobenzene                      (e) Benzenesulfonic acid              (f) Methoxybenzene
- 16.32** Rank the compounds in each group according to their reactivity toward electrophilic substitution.
- (a) Chlorobenzene, *o*-dichlorobenzene, benzene  
(b) *p*-Bromonitrobenzene, nitrobenzene, phenol  
(c) Fluorobenzene, benzaldehyde, *o*-xylene  
(d) Benzonitrile, *p*-methylbenzonitrile, *p*-methoxybenzonitrile
- 16.33** Predict the major monoalkylation products you would expect to obtain from reaction of the following substances with chloromethane and  $\text{AlCl}_3$ :
- (a) Bromobenzene                      (b) *m*-Bromophenol  
(c) *p*-Chloroaniline                      (d) 2,4-Dichloronitrobenzene  
(e) 2,4-Dichlorophenol                      (f) Benzoic acid  
(g) *p*-Methylbenzenesulfonic acid              (h) 2,5-Dibromotoluene
- 16.34** Name and draw the major product(s) of electrophilic chlorination of these substances:
- (a) *m*-Nitrophenol                      (b) *o*-Xylene  
(c) *p*-Nitrobenzoic acid                      (d) *p*-Bromobenzenesulfonic acid
- 16.35** Predict the major product(s) you would obtain from sulfonation of these compounds:
- (a) Fluorobenzene                      (b) *m*-Bromophenol                      (c) *m*-Dichlorobenzene  
(d) 2,4-Dibromophenol
- 16.36** Rank the following aromatic compounds in the expected order of their reactivity toward Friedel–Crafts alkylation. Which compounds are unreactive?
- (a) Bromobenzene                      (b) Toluene                      (c) Phenol  
(d) Aniline                      (e) Nitrobenzene                      (f) *p*-Bromotoluene
- 16.37** Suggest a reason for the observation that bromination of biphenyl occurs at ortho and para positions rather than at meta. Use resonance structures of the intermediates to explain your answers.



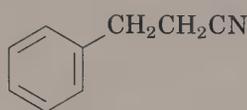
Biphenyl

- 16.38** At what position, and on what ring, would you expect nitration of 4-bromobiphenyl to occur?

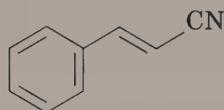


4-Bromobiphenyl

- 16.39 How do you explain the fact that electrophilic attack on 3-phenylpropanenitrile occurs at the ortho and para positions, whereas attack on 3-phenylpropenenitrile occurs at the meta position? Use resonance structures of the intermediates in your explanation.

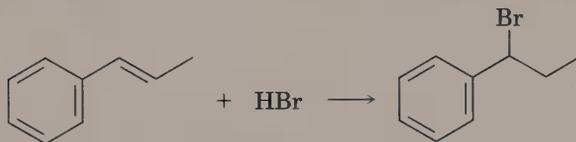


3-Phenylpropanenitrile

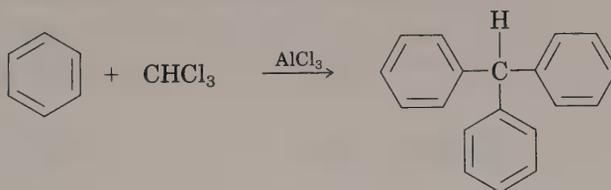


3-Phenylpropenenitrile

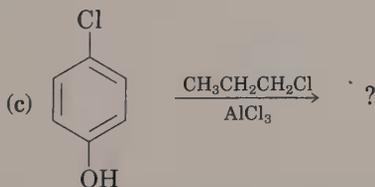
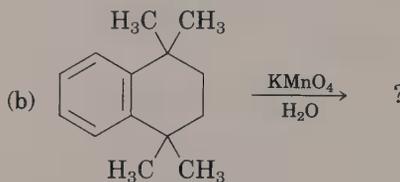
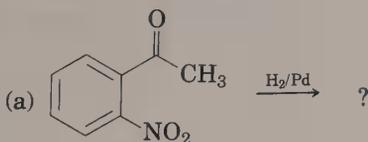
- 16.40 Addition of HBr to 1-phenylpropene yields (1-bromopropyl)benzene as the exclusive product. Propose a mechanism for the reaction, and explain why none of the other regioisomer is produced.



- 16.41 Triphenylmethane can be prepared by reaction of benzene and chloroform in the presence of  $\text{AlCl}_3$ . Propose a mechanism for the reaction.



- 16.42 What product(s) would you expect to obtain from these reactions?



- 16.43 How would you synthesize these substances, starting from benzene? Assume that ortho and para substitution products can be separated.

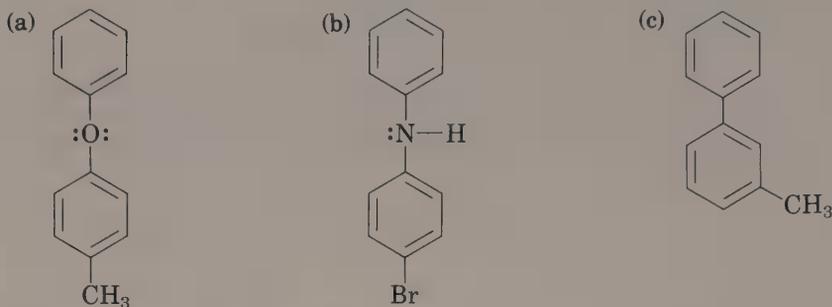
(a) *o*-Methylphenol

(b) 2,4,6-Trinitrophenol

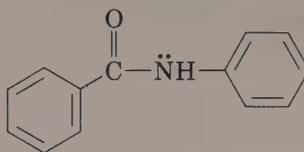
(c) 2,4,6-Trinitrobenzoic acid

(d) *m*-Bromoaniline

16.44 At what position, and on what ring, would you expect these substances to undergo electrophilic substitution?



16.45 At what position, and on what ring, would you expect bromination of benzamide to occur? Explain your answer by drawing resonance structures of the intermediates.



Benzamide

16.46 Would you expect the Friedel–Crafts reaction of benzene with optically active 2-chlorobutane to yield optically active or racemic product? Explain your answer.

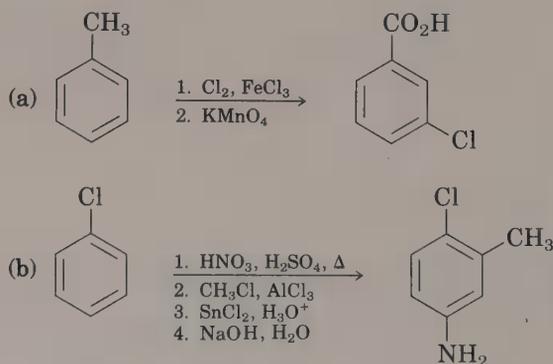
16.47 Starting with benzene as your only source of aromatic compounds, how would you synthesize these substances? Assume that you can separate ortho and para isomers if necessary.

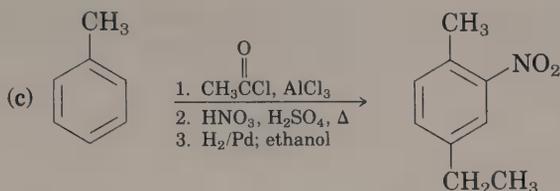
- (a) *p*-Chlorophenol  
(b) *m*-Bromonitrobenzene  
(c) *o*-Bromobenzenesulfonic acid  
(d) *m*-Chlorobenzenesulfonic acid

16.48 Starting with either benzene or toluene, how would you synthesize these materials? Assume that ortho and para isomers can be separated.

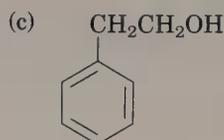
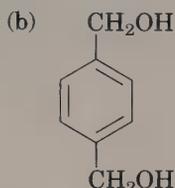
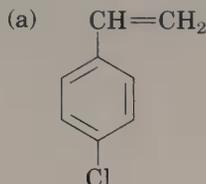
- (a) 2-Bromo-4-nitrotoluene  
(b) 1,3,5-Trinitrobenzene  
(c) 2,4,6-Tribromoaniline  
(d) 2-Chloro-4-methylphenol

16.49 As written, the following syntheses have certain flaws. What is wrong with each one?

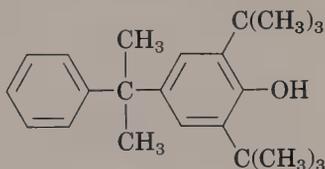




16.50 How would you synthesize these substances, starting from benzene?

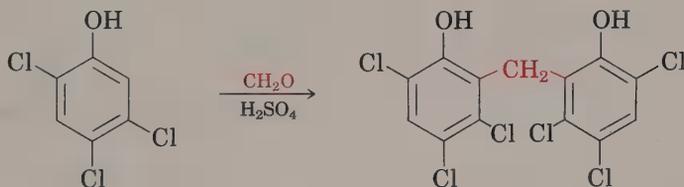


16.51 The compound MON-0585 is a nontoxic, biodegradable larvicide that is highly selective against mosquito larvae. How could you synthesize MON-0585 using only benzene as a source of the aromatic rings?



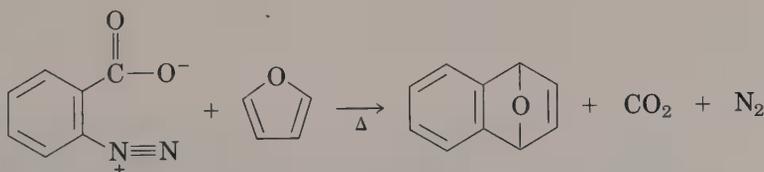
MON-0585

16.52 Hexachlorophene, a substance used in the manufacture of germicidal soaps, is prepared by reaction of 2,4,5-trichlorophenol with formaldehyde in the presence of concentrated sulfuric acid. Propose a mechanism to account for the reaction.



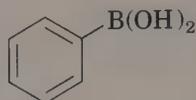
Hexachlorophene

16.53 When heated, benzenediazonium carboxylate decomposes to yield  $N_2$ ,  $CO_2$ , and a reactive organic substance that can't be isolated. When benzenediazonium carboxylate is heated in the presence of furan, the following reaction is observed:



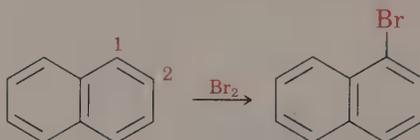
What intermediate is involved in this reaction? Propose a mechanism for formation of this intermediate.

- 16.54 Phenylboronic acid is nitrated to give 15% ortho substitution product and 85% meta. Account for the meta-directing effect of the  $-\text{B}(\text{OH})_2$  group.

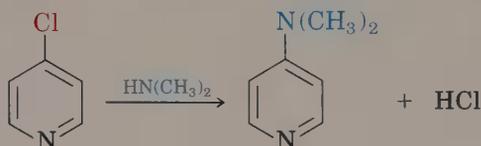


Phenylboronic acid

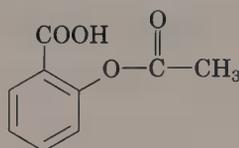
- 16.55 Draw resonance structures of the intermediate carbocations in the bromination of naphthalene, and account for the fact that naphthalene undergoes electrophilic attack at C1 rather than C2.



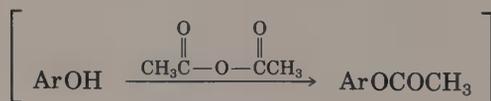
- 16.56 4-Chloropyridine undergoes reaction with dimethylamine to yield 4-dimethylaminopyridine. Propose a mechanism to account for this result.



- 16.57 How do you account for the fact that *p*-bromotoluene reacts with potassium amide to give a mixture of *m*- and *p*-methylaniline?
- 16.58 Propose a synthesis of aspirin (acetylsalicylic acid) starting from benzene. You will need to use an acetylation reaction at some point in your scheme.

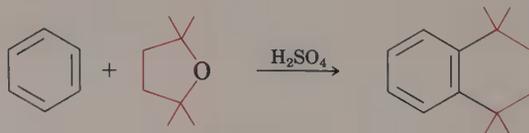


Aspirin

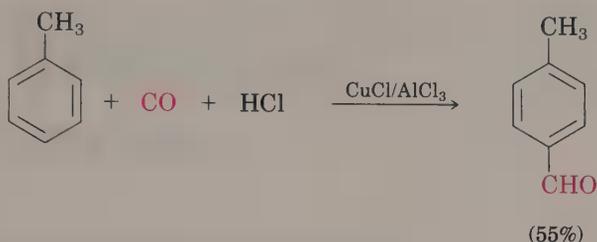


An acetylation reaction

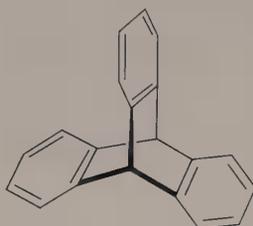
- 16.59 Propose a mechanism to account for the following reaction of benzene with 2,2,5,5-tetramethyltetrahydrofuran.



- 16.60 In the Gatterman–Koch reaction, a formyl group ( $-\text{CHO}$ ) is introduced directly onto a benzene ring. For example, reaction of toluene with carbon monoxide and  $\text{HCl}$  in the presence of mixed  $\text{CuCl}/\text{AlCl}_3$  gives *p*-methylbenzaldehyde in 55% yield. Propose a mechanism for this reaction.

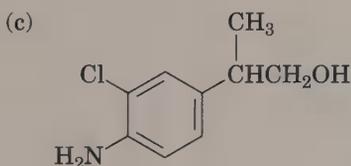
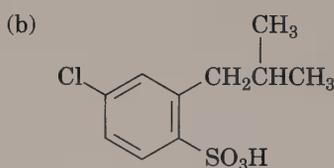
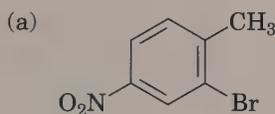


- 16.61** Triptycene is an unusual molecule that has been prepared by reaction of benzyne with anthracene. What kind of reaction is involved? Show the mechanism of the transformation.

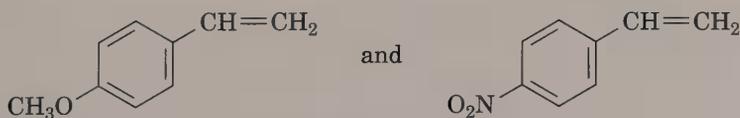


Triptycene

- 16.62** Treatment of *p*-*tert*-butylphenol with a strong acid such as  $\text{H}_2\text{SO}_4$  yields phenol and 2-methylpropene. Propose a mechanism for the reaction.
- 16.63** How would you synthesize these compounds from benzene? Assume that ortho and para isomers can be separated.

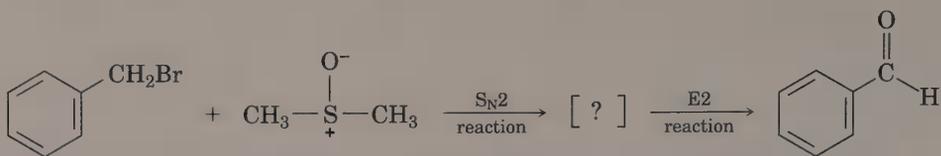


- 16.64** You know the mechanism of HBr addition to alkenes, and you know the effects of various substituent groups on aromatic substitution. Use this knowledge to predict which of the following two alkenes reacts faster with HBr. Explain your answer by drawing resonance structures of the carbocation intermediates.

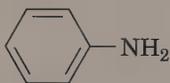


- 16.65** Draw a Fischer projection of (*R*)-2-phenylbutane, and predict the stereochemistry of its reaction with *N*-bromosuccinimide. Explain your answer.

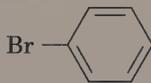
- 16.66 Benzyl bromide is converted into benzaldehyde by heating in dimethyl sulfoxide. Propose a structure for the intermediate, and show the mechanisms of the two steps in the reaction.



- 16.67 Use your knowledge of directing effects along with the following data to deduce the directions of the dipole moments in aniline and bromobenzene.



$$\mu = 1.53 \text{ D}$$



$$\mu = 1.52 \text{ D}$$



$$\mu = 2.91 \text{ D}$$

# Organic Reactions: A Brief Review

There's no question about it—learning organic chemistry means knowing a large number of reactions. The way to simplify the job, of course, is to organize the material. We said in Chapter 5 that organic reactions can be organized in two ways: by what kinds of reactions occur and by how they occur. Let's briefly review both organizational methods in light of what has been covered in the past several chapters.

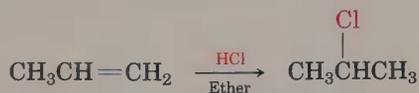
## I. A Summary of the Kinds of Organic Reactions

There are four important kinds of reactions: additions, eliminations, substitutions, and rearrangements. We've now seen examples of all four, as summarized in Review Tables 1–4.

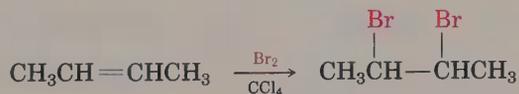
### REVIEW TABLE 1 Some Addition Reactions

#### 1. Additions to alkenes

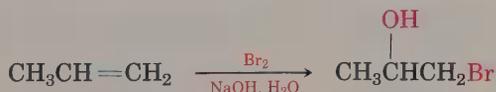
- a. Electrophilic addition of HX (X = Cl, Br, I; Sections 6.9 and 6.10)



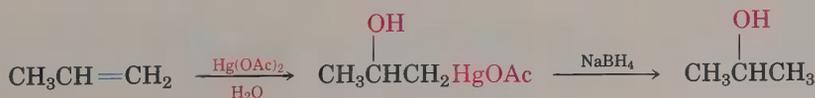
- b. Electrophilic addition of X<sub>2</sub> (X = Cl, Br; Section 7.2)



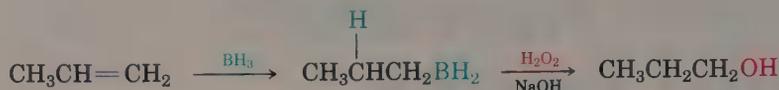
- c. Electrophilic addition of HO-X (X = Cl, Br, I; Section 7.3)



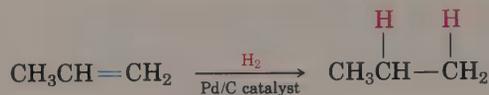
- d. Electrophilic addition of water by hydroxymercuration (Section 7.4)



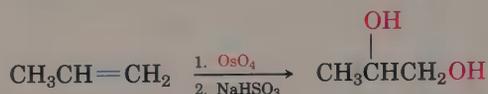
e. Addition of  $\text{BH}_3$  (hydroboration; Section 7.5)



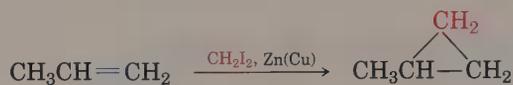
f. Catalytic addition of  $\text{H}_2$  (Section 7.7)



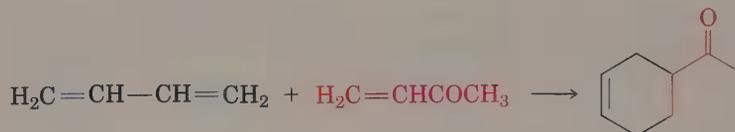
g. Hydroxylation with  $\text{OsO}_4$  (Section 7.8)



h. Addition of carbenoids: cyclopropane formation (Section 7.11)



i. Cycloaddition to dienes: Diels–Alder reaction (Sections 14.7 and 14.8)

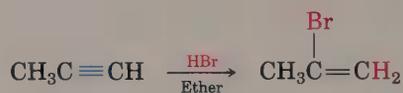


j. Radical addition of  $\text{HBr}$  (Section 7.6)

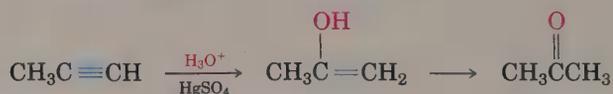


2. Additions to alkynes

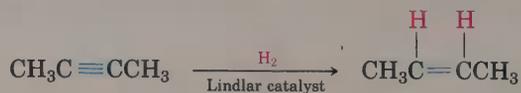
a. Electrophilic addition of  $\text{HX}$  ( $\text{X} = \text{Cl}, \text{Br}, \text{I}$ ; Section 8.4)



b. Electrophilic addition of  $\text{H}_2\text{O}$  (Section 8.5)



c. Addition of  $\text{H}_2$  (Section 8.7)

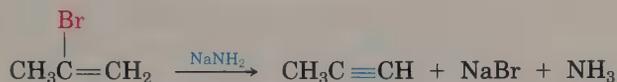


**REVIEW TABLE 2** Some Elimination Reactions

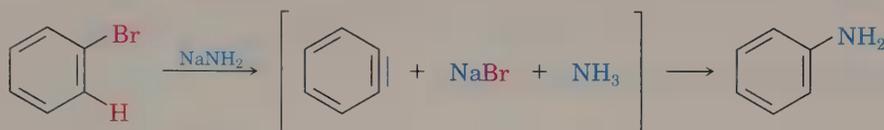
1. Dehydrohalogenation of alkyl halides (Sections 11.10 and 11.12)



2. Dehydrohalogenation of vinylic halides (Section 8.3)



3. Dehydrohalogenation of aryl halides: benzyne formation (Section 16.10)



4. Dehydration of alcohols (Section 7.1)

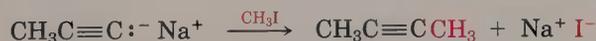
**REVIEW TABLE 3** Some Substitution Reactions

- 1.
- $S_N2$
- reactions of primary alkyl halides (Sections 11.2–11.5)

- a. General reaction

where  $\text{X} = \text{Cl}, \text{Br}, \text{I}, \text{OTos}$  $\text{:Nu}^- = \text{CH}_3\text{O}^-, \text{HO}^-, \text{CH}_3\text{S}^-, \text{HS}^-, \text{CN}^-, \text{CH}_3\text{COO}^-, \text{NH}_3, (\text{CH}_3)_3\text{N}$ , etc.

- b. Alkyne alkylation (Section 8.9)

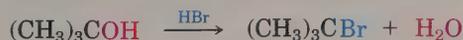


- 2.
- $S_N1$
- reactions of tertiary alkyl halides (Sections 11.8 and 11.9)

- a. General reaction

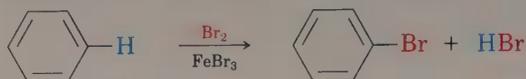


- b. Preparation of alkyl halides from alcohols (Section 10.8)

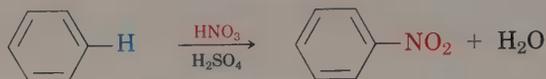


## 3. Electrophilic aromatic substitution (Sections 16.1–16.4)

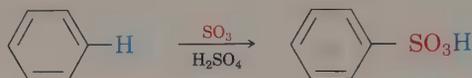
## a. Halogenation of aromatic compounds (Section 16.1)



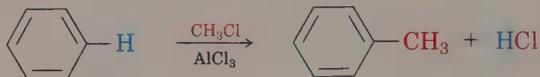
## b. Nitration of aromatic compounds (Section 16.2)



## c. Sulfonation of aromatic compounds (Section 16.2)



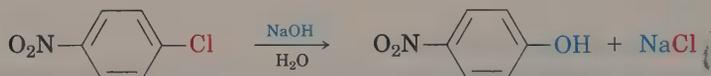
## d. Alkylation of aromatic rings (Section 16.3)



## e. Acylation of aromatic rings (Section 16.4)

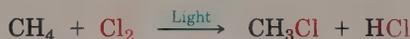


## 4. Nucleophilic aromatic substitution (Section 16.9)

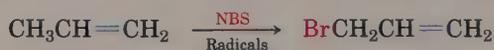


## 5. Radical substitution reactions

## a. Chlorination of methane (Section 10.4)



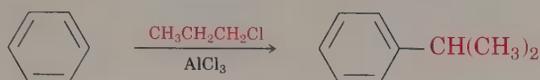
## b. NBS allylic bromination of alkenes (Section 10.5)

**REVIEW TABLE 4** Some Rearrangement Reactions

## 1. Carbocation rearrangement during electrophilic addition to alkenes (Section 6.13)



## 2. Carbocation rearrangement during Friedel–Crafts alkylation (Section 16.3)

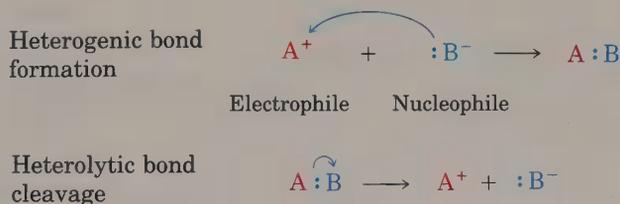


## II. A Summary of How Reactions Occur

The second method of organizing reactions is by how they occur—that is, by their *mechanisms*. We said in Chapter 5 that there are three fundamental reaction types: polar reactions, radical reactions, and pericyclic reactions. Having seen many different examples by now, let's see how this assertion stands up.

### A. POLAR REACTIONS

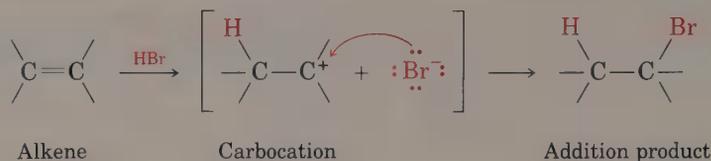
Polar reactions take place between electron-rich reagents (nucleophiles/Lewis bases) and electron-poor reagents (electrophiles/Lewis acids). These reactions are heterolytic processes and involve species with an even number of valence electrons. Bonds are made when a nucleophile donates an electron pair to an electrophile; bonds are broken when one product fragment leaves with an electron pair.



The polar reactions we've studied can be grouped into five general categories:

1. Electrophilic addition reactions
2. Elimination reactions
3. Nucleophilic alkyl substitution reactions
4. Electrophilic aromatic substitution reactions
5. Nucleophilic aromatic substitution reactions

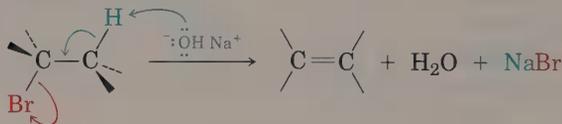
**1. Electrophilic Addition Reactions** (Sections 6.8–6.10) Alkenes react with electrophiles such as HBr to yield saturated addition products. The reaction occurs in two steps: The electrophile first adds to the alkene double bond to yield a carbocation intermediate that reacts further to yield the addition product.



Many of the addition reactions listed in Review Table 1 take place by an electrophilic addition mechanism. The electrophile may be  $\text{H}^+$  (HX addition; reactions 1a and 2a),  $\text{X}^+$  (halogen addition and halohydrin formation; reactions 1b and 1c), or  $\text{Hg}^{2+}$  (hydroxymercuration and alkyne hydration; reactions 1d and 2b), but the basic process is the same. The remaining addition reactions in Review Table 1 differ in that they occur without carbocation intermediates, but it's still convenient to group them together.

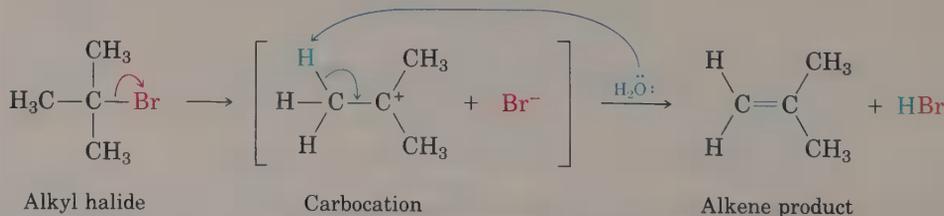
## 2. Elimination Reactions

*a. E2 reaction* (Sections 11.10 and 11.12) Alkyl halides undergo elimination of HX to yield alkenes on treatment with base. When a strong base such as hydroxide ion ( $\text{HO}^-$ ), alkoxide ion ( $\text{RO}^-$ ), or amide ion ( $\text{NH}_2^-$ ) is used, alkyl halides react by the E2 mechanism. E2 reactions occur in a single step involving removal by base of a neighboring hydrogen at the same time that the halide ion is leaving.



All of the elimination reactions listed in Review Table 2 occur by the same E2 mechanism. Though they appear different, the elimination of an alkyl halide to yield an alkene (reaction 1), the elimination of a vinylic halide to yield an alkyne (reaction 2), and the elimination of an aryl halide to yield a benzyne (reaction 3) are all E2 reactions.

*b. E1 reaction* (Section 11.14) Tertiary alkyl halides undergo elimination by the E1 mechanism in competition with  $\text{S}_{\text{N}}1$  substitution when a nonbasic nucleophile is used in a hydroxylic solvent. The reaction takes place in two steps: Spontaneous dissociation of the alkyl halide leads to a carbocation intermediate that then loses  $\text{H}^+$ .



## 3. Nucleophilic Alkyl Substitution Reactions

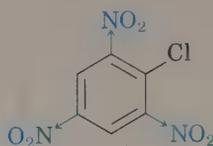
*a.  $\text{S}_{\text{N}}2$  reaction* (Sections 11.2–11.5) The nucleophilic alkyl substitution reaction is one of the most common reactions encountered in organic chem-



**5. Nucleophilic Aromatic Substitution Reactions** (Section 16.9) Nucleophilic aromatic substitution (reaction 4, Review Table 3) is a polar reaction that must be placed in a unique category, since it is not related to the other general processes we have studied. In this reaction, a nucleophile attacks an electrophilic aromatic ring. The ring is made electrophilic, and hence reactive, only when substituted by strong electron-withdrawing groups such as nitro, cyano, and carbonyl. Although this electrophilic behavior of the aromatic ring is a reversal of its normal reactivity, the process still involves reaction between an electrophile and a nucleophile, as with all polar reactions.



“Normal” aromatic ring;  
electron-rich and nucleophilic

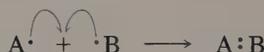


Electron-poor aromatic ring;  
electrophilic

## B. RADICAL REACTIONS

Radical reactions are homolytic processes that involve species with an odd number of electrons. Bonds are made when each reactant donates one electron, and bonds are broken when each product fragment leaves with one electron.

Homogenic bond  
formation



Homolytic bond  
cleavage



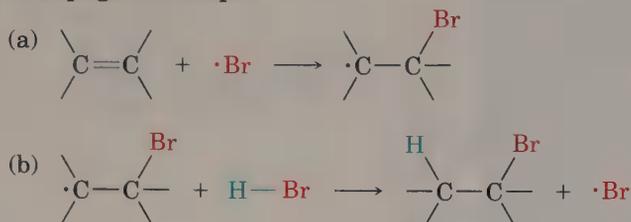
Since radical reactions are less common than polar reactions, we've seen only a few examples. Those we have studied can be classified as either radical addition reactions or radical substitution reactions. Radical additions such as the peroxide-catalyzed addition of HBr to alkenes (reaction 1j, Review Table 1) involve the addition of a radical to an unsaturated substrate. As we'll see in Chapter 31, this kind of reaction is extremely important in the preparation of polymers such as polystyrene.

Radical additions occur through three distinct kinds of steps, all of which involve odd-electron species: (1) initiation, (2) propagation, and (3) termination.

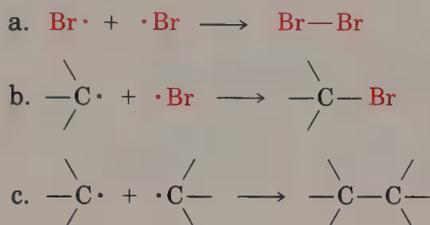
### 1. Initiation steps:



## 2. Propagation steps:



## 3. Termination steps:

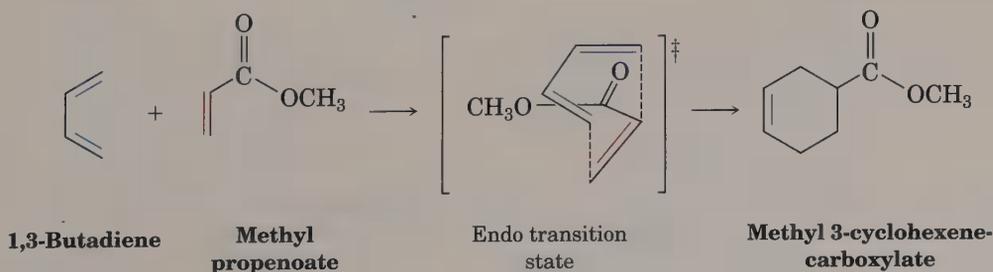


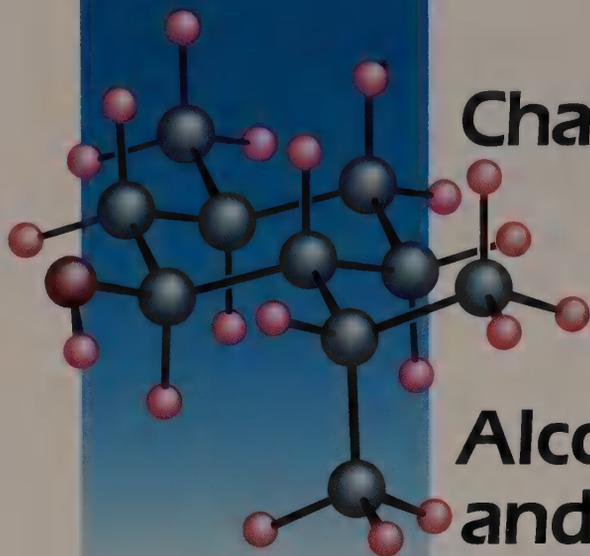
The reaction is initiated by thermal homolytic cleavage of a peroxide, which forms two radicals. These radicals abstract H· from HBr, yielding a Br· radical that then adds to the alkene, generating a new carbon radical and a carbon–bromine bond. Each reactant donates one electron to the new carbon–bromine bond. The reaction is completed by reaction of the carbon radical with HBr to yield neutral product and a bromine radical, which continues the chain. Note that the new carbon–hydrogen bond is also formed by donation of one electron from each reactant.

Radical substitution reactions, such as the light-induced chlorination of methane and the allylic bromination of alkenes with *N*-bromosuccinimide (reactions 5a and 5b, Review Table 3), are also common. The key feature of all these reactions is that one radical abstracts an atom from a neutral molecule, leaving a new radical.

## C. PERICYCLIC REACTIONS

Pericyclic reactions such as the addition of a carbene to an alkene and the Diels–Alder cycloaddition (reactions 1h and 1i, Review Table 1) involve neither radicals nor nucleophile–electrophile interactions. Rather, these processes take place in a single step by a reorganization of bonding electrons through a cyclic transition state. A fuller discussion of pericyclic reactions will be given in Chapter 30.

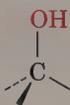




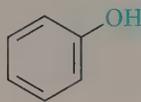
# Chapter 17

## Alcohols and Thiols

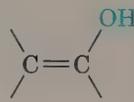
**Alcohols** are compounds that have hydroxyl groups bonded to saturated,  $sp^3$ -hybridized carbon atoms. This definition purposely excludes phenols (hydroxyl groups bonded to an aromatic ring) and enols (hydroxyl groups bonded to a vinylic carbon), because the chemistry of the three types of compounds is sometimes different. Alcohols can be thought of as organic derivatives of water in which one of the water hydrogens is replaced by an organic group:  $H-O-H$  versus  $R-O-H$ .



An alcohol



A phenol

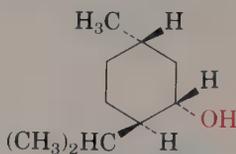


An enol

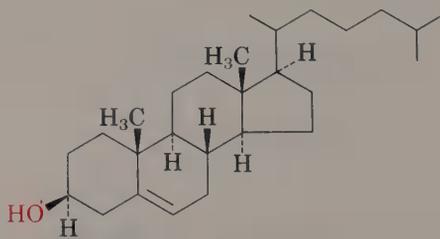
Alcohols occur widely in nature and have a great many industrial and pharmaceutical applications. Ethanol, for instance, is one of the simplest yet best known of all organic substances, finding use as a fuel additive, an industrial solvent, and a beverage; menthol, an alcohol isolated from peppermint oil, is widely used as a flavoring and perfumery agent; and cholesterol, a complicated-looking steroidal alcohol, has been implicated as a causative agent in heart disease.



Ethanol



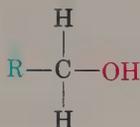
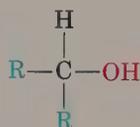
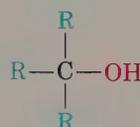
Menthol



Cholesterol (a steroid)

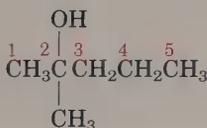
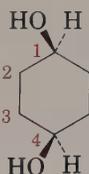
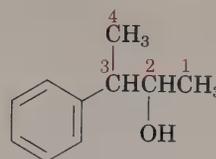
## 17.1 Nomenclature of Alcohols

Alcohols are classified as primary ( $1^\circ$ ), secondary ( $2^\circ$ ), or tertiary ( $3^\circ$ ), depending on the number of carbon substituents bonded to the hydroxyl-bearing carbon.

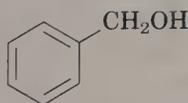
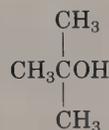
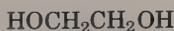
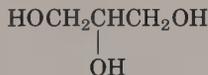
A primary alcohol ( $1^\circ$ )A secondary alcohol ( $2^\circ$ )A tertiary alcohol ( $3^\circ$ )

Simple alcohols are named by the IUPAC system as derivatives of the parent alkane, using the suffix *-ol*:

1. Select the longest carbon chain *containing the hydroxyl group*, and derive the parent name by replacing the *-e* ending of the corresponding alkane with *-ol*.
2. Number the alkane chain beginning at the end nearer the hydroxyl group.
3. Number the substituents according to their position on the chain, and write the name listing the substituents in alphabetical order.

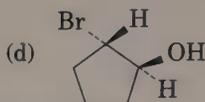
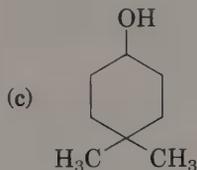
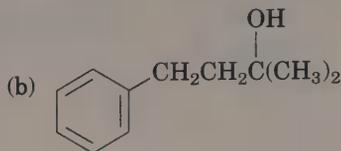
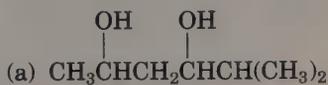
**2-Methyl-2-pentanol****cis-1,4-Cyclohexanediol****3-Phenyl-2-butanol**

Certain simple and widely occurring alcohols have common names that are accepted by IUPAC. For example:

**Benzyl alcohol**  
(phenylmethanol)**Allyl alcohol**  
(2-propen-1-ol)**tert-Butyl alcohol**  
(2-methyl-2-propanol)**Ethylene glycol**  
(1,2-ethanediol)**Glycerol**  
(1,2,3-propanetriol)

## PROBLEM.....

17.1 Give IUPAC names for these compounds:



## PROBLEM.....

17.2 Draw structures corresponding to these IUPAC names:

(a) 2-Ethyl-2-buten-1-ol

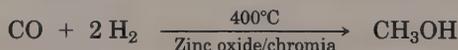
(b) 3-Cyclohexen-1-ol

(c) *trans*-3-Chlorocycloheptanol

(d) 1,4-Pentanediol

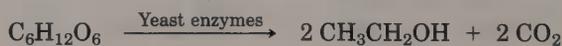
## 17.2 Sources and Uses of Simple Alcohols

Methanol and ethanol are two of the most important of all industrial chemicals. Prior to the development of the modern chemical industry, methanol was prepared by heating wood in the absence of air and thus came to be called *wood alcohol*. Today, approximately 1.2 billion gallons of methanol are manufactured each year in the United States by catalytic reduction of carbon monoxide with hydrogen gas.



Methanol is toxic to humans, causing blindness in low doses and death in larger amounts. Industrially, it is used both as a solvent and as a starting material for production of formaldehyde,  $\text{CH}_2\text{O}$ , and acetic acid,  $\text{CH}_3\text{COOH}$ .

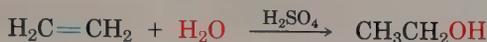
Ethanol is one of the oldest known pure organic chemicals. Its production by fermentation of grains and sugars, and its subsequent purification by distillation, go back at least as far as the 12th century A.D. Fermentation is carried out by adding yeast to an aqueous sugar solution where enzymes in the yeast break down carbohydrates into ethanol and  $\text{CO}_2$ .



Glucose

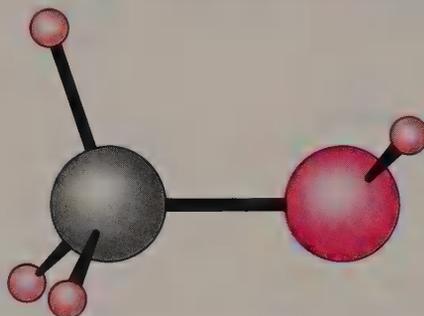
Only about 5% of the ethanol produced industrially comes from fermentation, although that figure may well change as demand for use in automobile fuel increases. Most ethanol is currently obtained by acid-catalyzed hydration of ethylene (Section 7.4). Nearly 300 million gallons of

ethanol a year are produced in the United States for use as a solvent or as a chemical intermediate in other industrial reactions.



## 17.3 Properties of Alcohols: Hydrogen Bonding

As mentioned earlier, alcohols can be thought of as organic derivatives of water in which one of the hydrogens has been replaced by an organic group. As such, alcohols have nearly the same geometry as water. The R–O–H bond angle has an approximately tetrahedral value ( $109^\circ$  in methanol, for example), and the oxygen atom is  $sp^3$  hybridized.

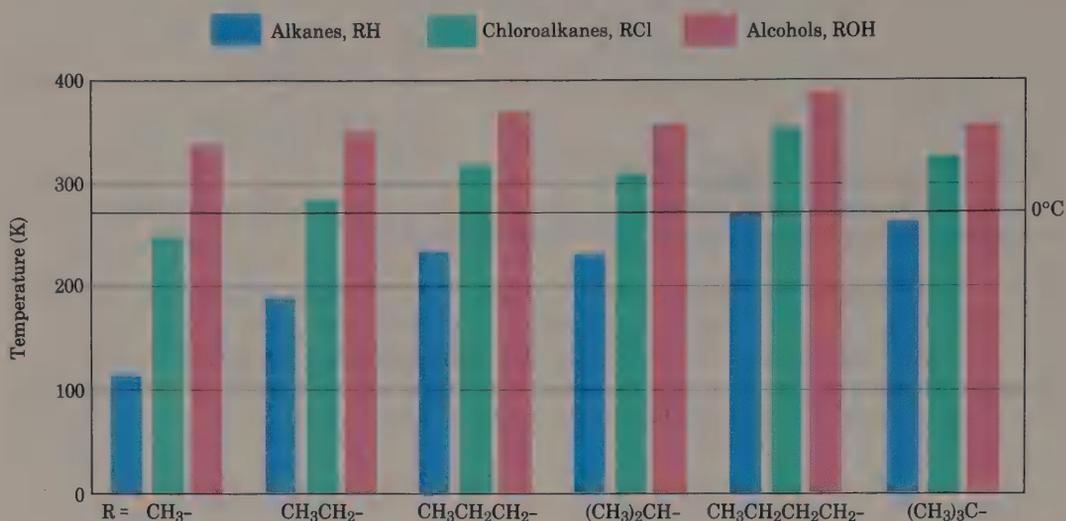


Methanol,  $\text{CH}_3\text{OH}$

Alcohols are quite different from the hydrocarbons and alkyl halides studied thus far. Not only is their chemistry much richer, but their physical properties are different as well. Table 17.1 and Figure 17.1, which provide a comparison of the boiling points of some simple alcohols, alkanes, and chloroalkanes, show that alcohols have much higher boiling points. For example, 1-propanol (mol wt = 60), butane (mol wt = 58), and chloroethane (mol wt = 65) are close in formula weight, yet 1-propanol boils at  $97^\circ\text{C}$ , compared to  $-0.5^\circ\text{C}$  for the alkane and  $12.5^\circ\text{C}$  for the chloroalkane.

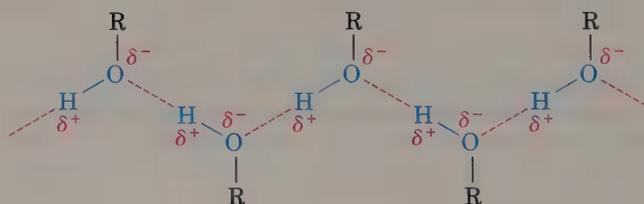
**TABLE 17.1** Boiling Points of Alkanes, Chloroalkanes, and Alcohols ( $^\circ\text{C}$ )

Alkyl group, R	Alkane, R—H	Chloroalkane, R—Cl	Alcohol, R—OH
$\text{CH}_3-$	-162	-24	64.5
$\text{CH}_3\text{CH}_2-$	-88.5	12.5	78.3
$\text{CH}_3\text{CH}_2\text{CH}_2-$	-42	46.6	97
$(\text{CH}_3)_2\text{CH}-$	-42	36.5	82.5
$\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2-$	-0.5	83.5	117
$(\text{CH}_3)_3\text{C}-$	-12	51	83



**Figure 17.1** A comparison of boiling points for some alkanes, chloroalkanes, and alcohols.

The reason for their high boiling points is that alcohols, like water, are highly associated in solution because of the formation of **hydrogen bonds**. The positively polarized  $\text{-OH}$  hydrogen atom from one molecule forms a weak hydrogen bond to the negatively polarized oxygen atom of another molecule (Figure 17.2). Although hydrogen bonds have a strength of only about 5 kcal/mol (20 kJ/mol) versus 103 kcal/mol (431 kJ/mol) for a typical O–H covalent bond, the presence of many hydrogen bonds means that extra energy must be added to break them during the boiling process.



**Figure 17.2** Hydrogen bonding in alcohols.

**PROBLEM** .....

- 17.3** The following data for three isomeric four-carbon alcohols show that there is a decrease in boiling point with increasing substitution:

1-Butanol, bp 117.5°C

2-Butanol, bp 99.5°C

2-Methyl-2-propanol, bp 82.2°C

Propose an explanation to account for this trend.

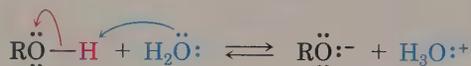
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## 17.4 Properties of Alcohols: Acidity and Basicity

Like water, alcohols are both weakly acidic and weakly basic. As weak Lewis bases, alcohols are reversibly protonated by acids to yield oxonium ions,  $\text{ROH}_2^+$ . As you might expect, protonated alcohols are much more reactive than neutral alcohols toward nucleophiles because they bear a positive charge.



As weak acids, alcohols act as proton donors. In dilute aqueous solution, alcohols dissociate to a slight extent by donating a proton to water.



In our earlier discussion of acidity (Sections 2.6 and 2.7), we said that the strength of any acid HA in water can be defined by the expressions

$$K_a = \frac{[\text{A}^-][\text{H}_3\text{O}^+]}{[\text{HA}]} \quad \text{and} \quad \text{p}K_a = -\log K_a$$

where  $K_a$  is the acidity constant

Compounds with a smaller  $K_a$  (or larger  $\text{p}K_a$ ) are weakly acidic, whereas compounds with a larger  $K_a$  (or smaller  $\text{p}K_a$ ) are more strongly acidic.

The data presented in Table 17.2 show that simple alcohols are about as acidic as water but that structure plays a significant role in determining the exact acidity of a compound. For example, methanol and ethanol are similar to water in acidity, while *tert*-butyl alcohol is slightly less acidic.

The effect of alkyl substitution on acidity is due primarily to solvation. Water is able to surround the sterically accessible oxygen atom of unhindered alcohols such as methanol and to stabilize the alkoxide ion by solvation, thus favoring its formation. Hindered alkoxide ions such as *tert*-butoxide, however, prevent solvation by their bulk and are therefore less stabilized.

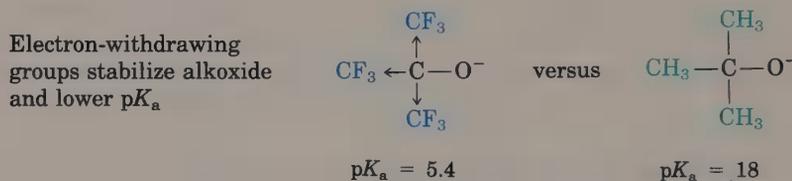
Inductive effects (Section 16.5) are also important in determining alcohol acidities. For example, electron-withdrawing halogen substituents stabilize an alkoxide anion by helping to spread out the charge over a larger

TABLE 17.2 Acidity Constants of Some Alcohols

Alcohol	$pK_a$	
$(\text{CH}_3)_3\text{COH}$	18.00	Weaker acid
$\text{CH}_3\text{CH}_2\text{OH}$	16.00	
$\text{HOH}$ (water) <sup>a</sup>	(15.74)	
$\text{CH}_3\text{OH}$	15.54	
$\text{CF}_3\text{CH}_2\text{OH}$	12.43	
$(\text{CF}_3)_3\text{COH}$	5.4	
$\text{HCl}$ (hydrochloric acid) <sup>a</sup>	(-7.00)	Stronger acid

<sup>a</sup>Values for water and hydrochloric acid are shown for reference.

area, thus making the alcohol more acidic. This inductive effect can be observed by comparing the acidities of ethanol ( $pK_a = 16$ ) and 2,2,2-trifluoroethanol ( $pK_a = 12.43$ ), or of *tert*-butyl alcohol ( $pK_a = 18$ ) and perfluoro-2-methyl-2-propanol ( $pK_a = 5.4$ ).



Since alcohols are much weaker than carboxylic acids or mineral acids, they don't react with weak bases such as amines, bicarbonate ion, or metal hydroxides. Alcohols do, however, react with alkali metals and with strong bases such as sodium hydride (NaH), sodium amide (NaNH<sub>2</sub>), alkyllithium reagents (RLi), and Grignard reagents (RMgX). The metal salts of alcohols are themselves strong bases that are frequently used as reagents in organic chemistry.



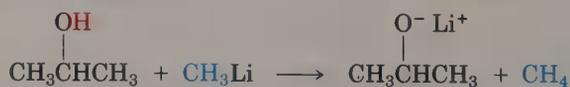
Methanol

Sodium methoxide



Ethanol

Sodium ethoxide

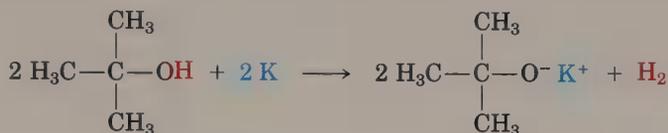


Isopropyl alcohol

Lithium isopropoxide



Cyclohexanol

Bromomagnesium  
cyclohexoxide*tert*-Butyl alcoholPotassium *tert*-butoxide

## PROBLEM .....

- 17.4 Which alcohol do you think is more acidic, benzyl alcohol or *p*-nitrobenzyl alcohol? Explain.
- .....

## 17.5 Preparation of Alcohols

Alcohols occupy a central position in organic chemistry. They can be prepared from a variety of functional groups (alkenes, alkyl halides, ketones, esters, and aldehydes, among others), and they can be transformed into a wide assortment of compound types (Figure 17.3).

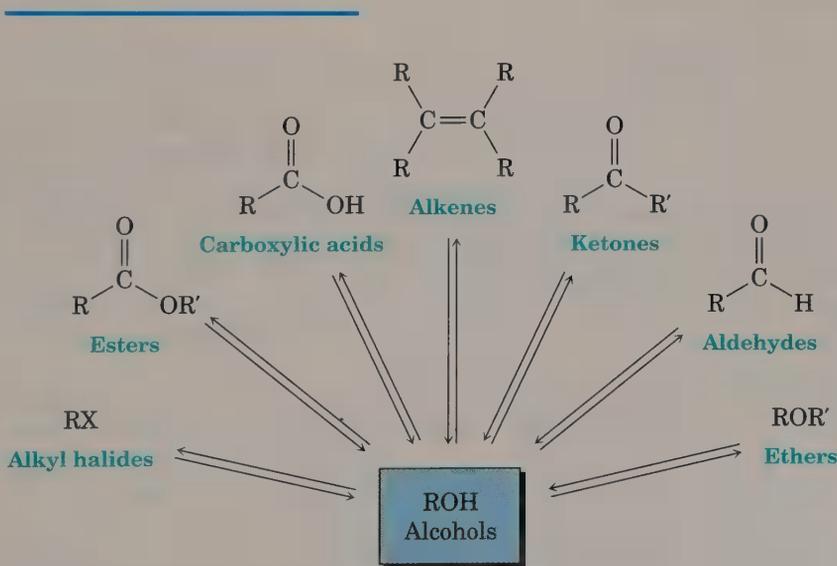
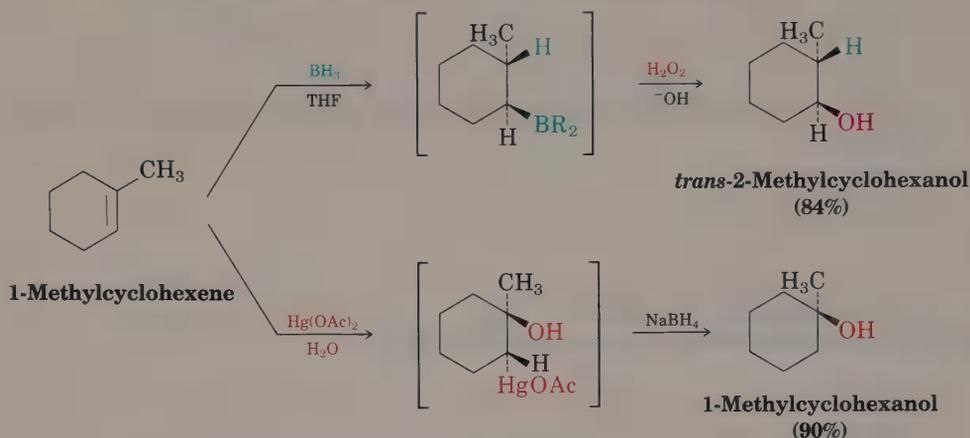


Figure 17.3 The central position of alcohols in organic chemistry.

Let's review briefly some of the methods of alcohol preparation we've already seen. Alcohols can be prepared by hydration of alkenes. Since the direct hydration of alkenes with aqueous acid is generally a poor reaction in the laboratory, two indirect methods are commonly used. A choice between the two is made depending on the product desired. Hydroboration–oxidation (Section 7.5) yields the product of syn non-Markovnikov hydration, whereas oxymercuration–reduction (Section 7.4) yields the product of Markovnikov hydration. Both reactions are mild and are generally applicable to most alkenes (Figure 17.4).

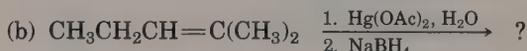
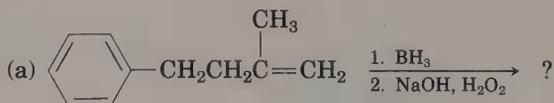


**Figure 17.4** Two complementary methods for the hydration of an alkene to yield an alcohol.

1,2-Diols can be prepared by direct hydroxylation of an alkene with osmium tetroxide followed by sodium bisulfite reduction (Section 7.8). The reaction takes place readily and is routinely used for the preparation of *cis* 1,2-diols (Figure 17.5). We'll see in the next chapter that 1,2-diols can also be prepared by acid-catalyzed hydrolysis of *epoxides*—compounds with a three-membered oxygen-containing ring. This method of epoxide opening is complementary to direct hydroxylation since it yields a *trans* 1,2-diol.

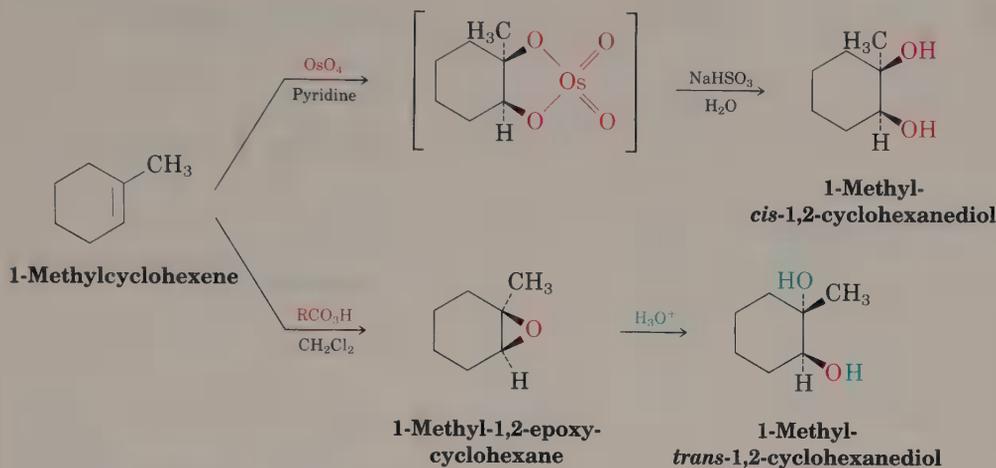
**PROBLEM** .....

**17.5** Predict the products of these reactions:



(c) Reaction of *cis*-5-decene with  $\text{OsO}_4$ , followed by  $\text{NaHSO}_3$  reduction. Be sure to indicate the stereochemistry of the product.

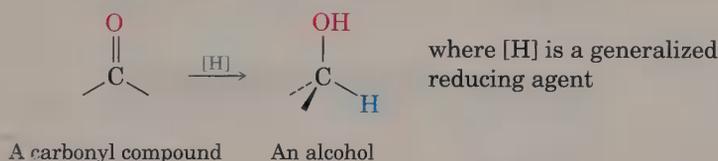
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**Figure 17.5** Two complementary methods for the preparation of 1,2-diols.

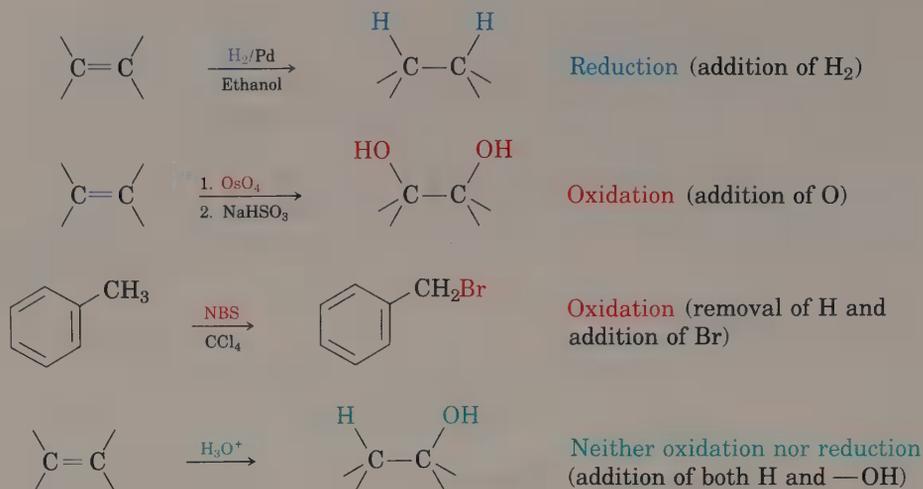
## 17.6 Alcohols from Reduction of Carbonyl Compounds

The most valuable method for preparing alcohols is by *reduction* of a carbonyl compound:

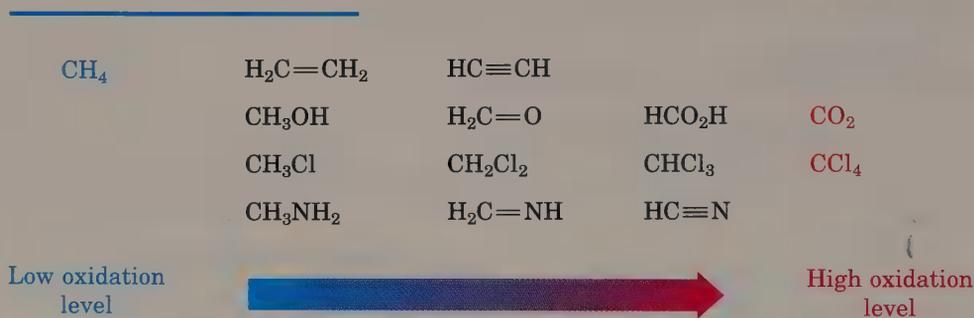


In inorganic chemistry, reduction is defined as the gain of electrons, and oxidation is defined as the loss of electrons. In organic chemistry, however, it's often difficult to decide whether an atom gains or loses electrons during a reaction. Thus, the terms oxidation and reduction have less precise meanings. For our purposes, an **organic reduction** is a reaction that either *increases* the hydrogen content or *decreases* the oxygen, nitrogen, or halogen content of a molecule. Conversely, an **organic oxidation** is a reaction that either decreases the hydrogen content or increases the oxygen, nitrogen, or halogen content of a molecule.

For example, catalytic hydrogenation of an alkene is clearly a reduction, since two hydrogens are added to the starting material. Hydroxylation of an alkene with  $\text{OsO}_4$ , however, is clearly an oxidation since oxygen is added. Similarly, NBS allylic bromination of an alkene is an oxidation since a hydrogen is replaced by a halogen. Hydration of an alkene is neither an oxidation nor a reduction, since both hydrogen and oxygen are added to the alkene at the same time.



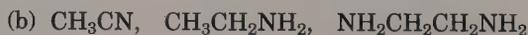
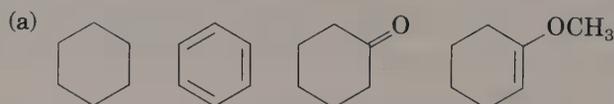
A list of functional groups of increasing oxidation level is shown in Figure 17.6. Any reaction that converts a functional group from a lower level to a higher level is an oxidation; any reaction that converts a functional group from a higher level to a lower level is a reduction; and any reaction that doesn't change the level is neither an oxidation nor a reduction.



**Figure 17.6** Oxidation levels of some common functional groups.

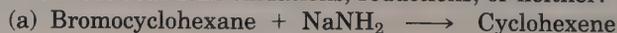
**PROBLEM** .....

**17.6** Rank the following series of compounds in order of increasing oxidation level:



**PROBLEM** .....

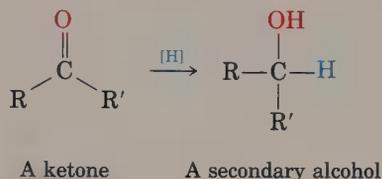
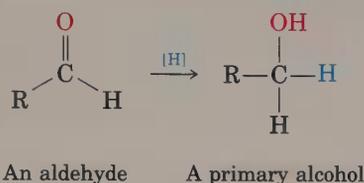
**17.7** Are these reactions oxidations, reductions, or neither?



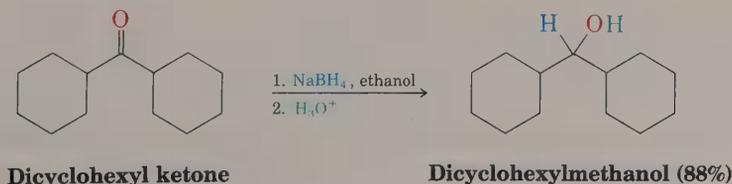
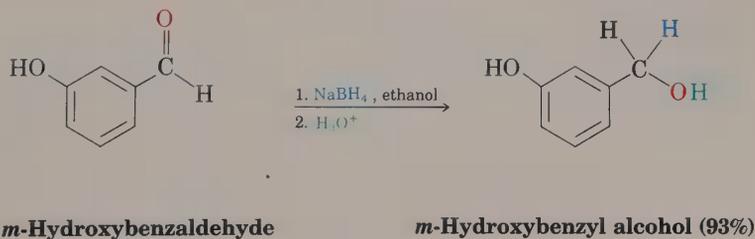
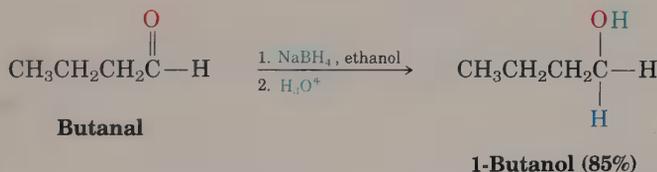
- (b) Benzene +  $\text{Cl}_2/\text{FeCl}_3 \longrightarrow$  Chlorobenzene  
 (c) 1-Bromobutane + Mg, then  $\text{H}_3\text{O}^+ \longrightarrow$  Butane  
 (d) Benzene +  $\text{CH}_3\text{Cl}/\text{AlCl}_3 \longrightarrow$  Toluene

## REDUCTION OF ALDEHYDES AND KETONES

Aldehydes and ketones are easily reduced to yield alcohols. Aldehydes are converted into primary alcohols, and ketones are converted into secondary alcohols on reduction.

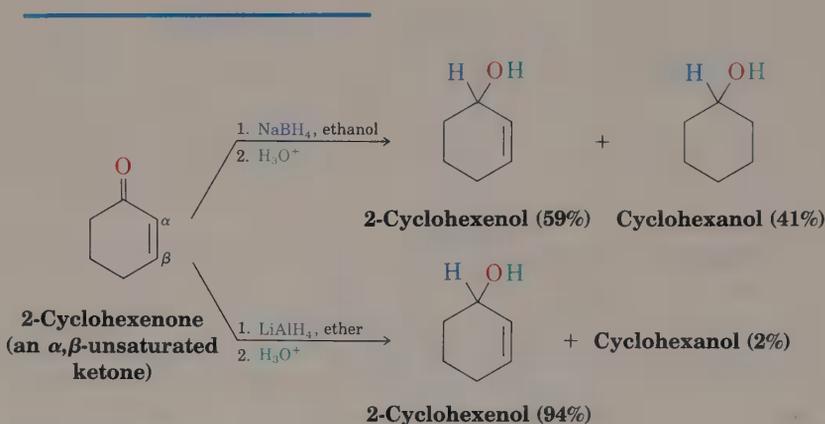


Many reagents are available for reducing ketones and aldehydes to alcohols, but sodium borohydride,  $\text{NaBH}_4$ , is usually chosen because of its safety and ease of handling. Sodium borohydride is a white, crystalline solid that can be safely handled and weighed in the open atmosphere and that can be used either in water or in alcohol solution. High yields of alcohol products are usually obtained, as the following examples indicate:



Another reducing agent that is sometimes used for reduction of ketones and aldehydes is lithium aluminum hydride,  $\text{LiAlH}_4$ , a white powder soluble in ether and tetrahydrofuran. Although more powerful and reactive than  $\text{NaBH}_4$ ,  $\text{LiAlH}_4$  is also far more dangerous and should be handled only by skilled persons. It reacts violently with water, decomposes when heated above  $120^\circ\text{C}$ , and has even been known to explode when being ground with a mortar and pestle. Despite these drawbacks,  $\text{LiAlH}_4$  is an extremely valuable reagent that is used daily in thousands of laboratories.

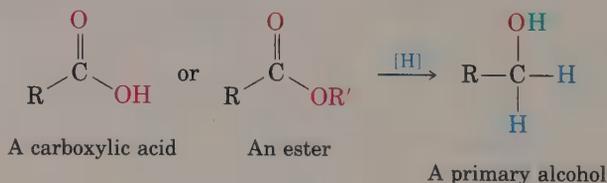
$\text{LiAlH}_4$  is particularly useful for reducing  $\alpha,\beta$ -unsaturated ketones (ketones conjugated with carbon-carbon double bonds). Although  $\alpha,\beta$ -unsaturated ketones often undergo overreduction with  $\text{NaBH}_4$  to give a mixture of both unsaturated alcohol and saturated alcohol, clean reduction to the allylic alcohol occurs with  $\text{LiAlH}_4$ . Thus, 2-cyclohexenone gives a 59:41 mixture of two products when  $\text{NaBH}_4$  is used as a reducing agent but gives largely 2-cyclohexenol when  $\text{LiAlH}_4$  is used (Figure 17.7).



**Figure 17.7** The reduction of 2-cyclohexenone with  $\text{NaBH}_4$  and  $\text{LiAlH}_4$ .

## REDUCTION OF ESTERS AND CARBOXYLIC ACIDS

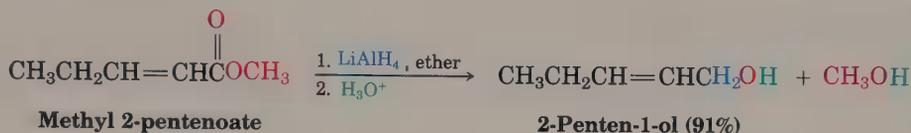
Esters and carboxylic acids can be reduced to give primary alcohols:



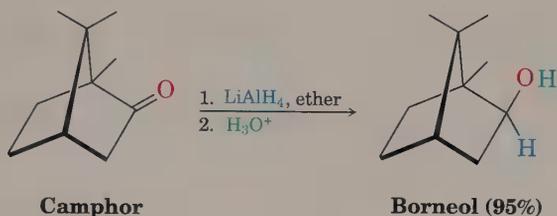
These reactions are more difficult than the analogous reductions of aldehydes and ketones. For example, sodium borohydride only slowly reduces esters and does not reduce acids at all. Ester and carboxylic acid reductions are therefore usually carried out with lithium aluminum hydride. All carbonyl groups, including esters, acids, ketones, and aldehydes, are reduced

by  $\text{LiAlH}_4$  in high yield, as some of the following examples indicate. Note that *one* hydrogen atom is delivered to the carbonyl carbon atom during reductions of ketones and aldehydes, but that *two* hydrogens become bonded to the carbonyl carbon during ester and carboxylic acid reductions.

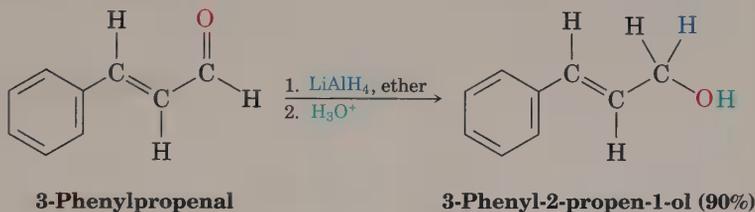
### Ester Reduction



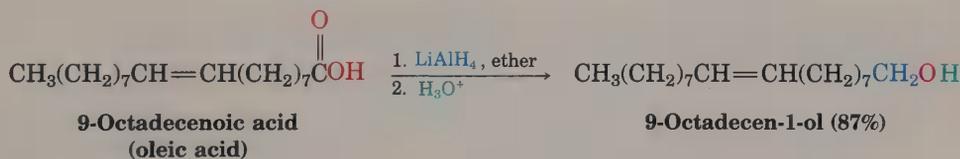
### Ketone Reduction



### Aldehyde Reduction

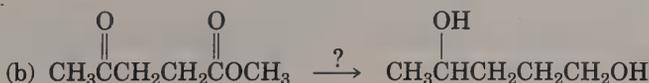
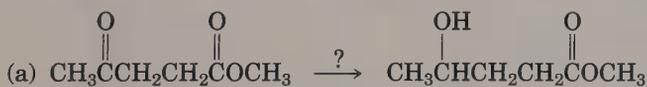


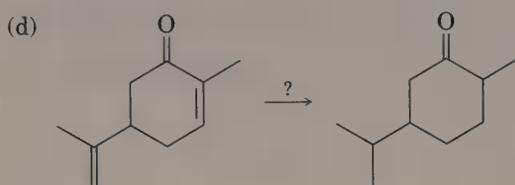
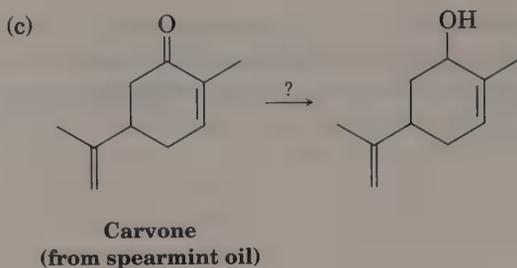
### Carboxylic Acid Reduction



### PROBLEM

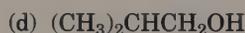
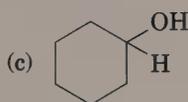
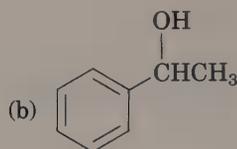
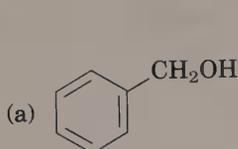
17.8 What reagent would you use to accomplish each of these reactions?





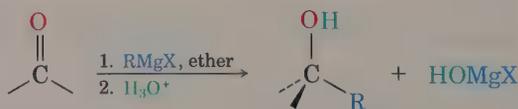
PROBLEM.....

- 17.9 What carbonyl compounds give the following alcohols on reduction with  $\text{LiAlH}_4$ ? Show all possibilities.



## 17.7 Alcohols from Addition of Grignard Reagents to Carbonyl Groups

Grignard reagents,  $\text{RMgX}$ , react with carbonyl compounds to yield alcohols in much the same way that hydride reagents do. The result is a useful and general method of alcohol synthesis.



We saw in Section 10.9 that alkyl, aryl, and vinylic halides react with magnesium in ether or tetrahydrofuran solution to generate Grignard reagents,  $\text{RMgX}$ .

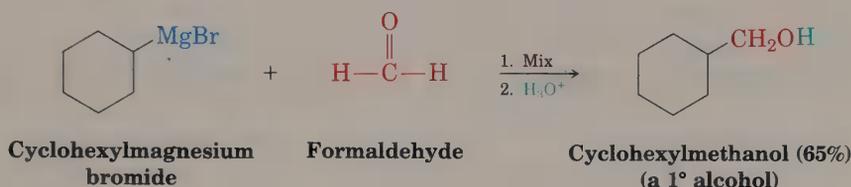


A Grignard reagent

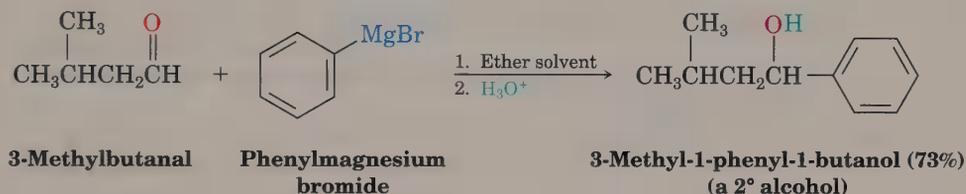
where R = 1°, 2°, or 3° alkyl, aryl, or vinylic

X = Cl, Br, or I

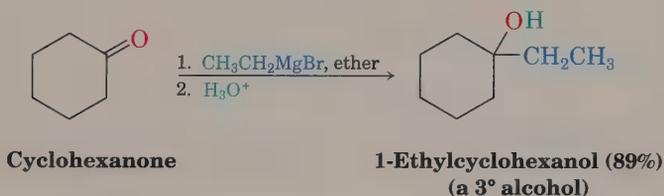
A large number of alcohol products can be obtained from Grignard reactions, depending on the reagents used. For example, Grignard reagents react with formaldehyde,  $\text{H}_2\text{C}=\text{O}$ , to give primary alcohols:



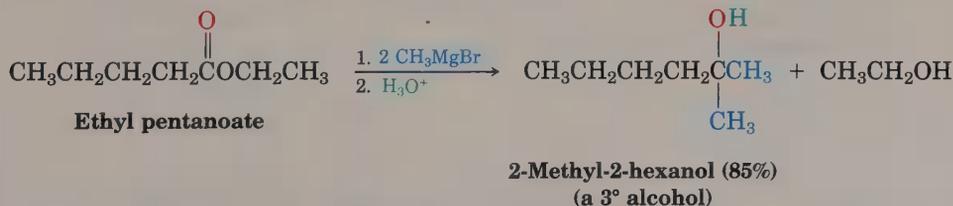
Aldehydes react with Grignard reagents to give secondary alcohols:



Ketones react similarly to yield tertiary alcohols:

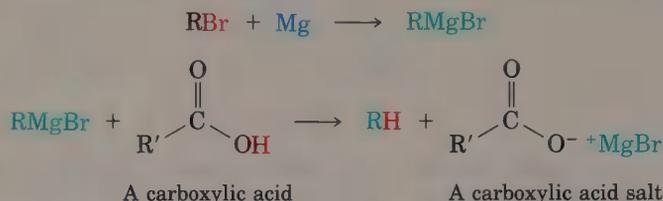


Esters react with Grignard reagents to yield tertiary alcohols in which two of the substituents bonded to the hydroxyl-bearing carbon have come from the Grignard reagent (just as  $\text{LiAlH}_4$  reduction of esters adds *two* hydrogens). For example:



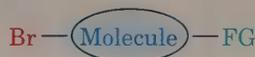
Carboxylic acids don't give addition products with Grignard reagents because the acidic carboxyl proton reacts with the Grignard reagent to

produce a hydrocarbon and the magnesium salt of the acid. We saw this reaction in Section 10.9 as a means of reducing alkyl halides to alkanes.

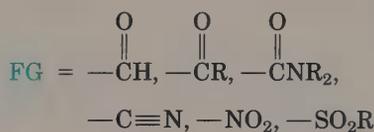


The Grignard reaction, though useful, has severe limitations. The major problem is that a Grignard reagent can't be prepared from an organohalide if there are other reactive functional groups in the same molecule. For example, a compound that is both an alkyl halide and a ketone won't form a Grignard reagent; instead, it reacts with itself. Similarly, a compound that is both an alkyl halide and a carboxylic acid, an alcohol, or an amine can't form a Grignard reagent because the acidic  $\text{RCO}_2\text{H}$ ,  $\text{ROH}$ , or  $\text{RNH}_2$  protons present in the same molecule react with the basic Grignard reagent as rapidly as it forms.

Generally speaking, Grignard reagents can't be prepared from alkyl halides that also contain these functional groups (FG):



where FG =  $-\text{OH}$ ,  $-\text{NH}$ ,  $-\text{SH}$ ,  $-\text{COOH}$



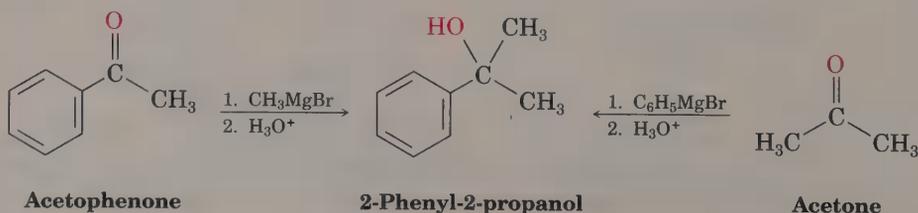
The Grignard reagent is protonated by these groups.

The Grignard reagent adds to these groups.

#### PRACTICE PROBLEM.....

How could you use the addition of a Grignard reagent to a ketone to synthesize 2-phenyl-2-propanol?

**Solution** First, draw the structure of the product and identify the groups bonded to the alcohol carbon atom. In the present instance, there are two methyl groups ( $-\text{CH}_3$ ) and one phenyl ( $-\text{C}_6\text{H}_5$ ). One of the three will have come from a Grignard reagent, and the remaining two will have come from a ketone. Thus, the possibilities are addition of methylmagnesium bromide to acetophenone and addition of phenylmagnesium bromide to acetone:



## PROBLEM .....

- 17.10 Show the products obtained from addition of methylmagnesium bromide to these compounds:  
 (a) Cyclopentanone      (b) Benzophenone (diphenyl ketone)      (c) 3-Hexanone

## PROBLEM .....

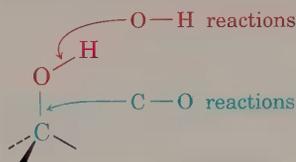
- 17.11 How could you use a Grignard addition reaction to prepare these alcohols?  
 (a) 2-Methyl-2-propanol      (b) 1-Methylcyclohexanol      (c) 3-Methyl-3-pentanol  
 (d) 2-Phenyl-2-butanol      (e) Benzyl alcohol

## PROBLEM .....

- 17.12 How can you explain the observation that treatment of 4-hydroxycyclohexanone with one equivalent of methylmagnesium bromide yields none of the expected addition product, whereas treatment with an excess of Grignard reagent leads to a good yield of 1-methyl-1,4-cyclohexanediol?

## 17.8 Reactions of Alcohols

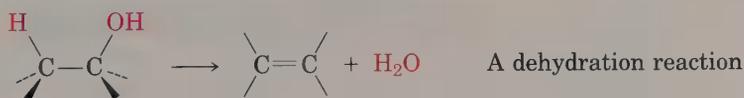
Reactions of alcohols can be divided into two groups—those that occur at the C–O bond and those that occur at the O–H bond:



Let's begin to look at reactions of both types by reviewing some of the alcohol reactions seen in previous chapters.

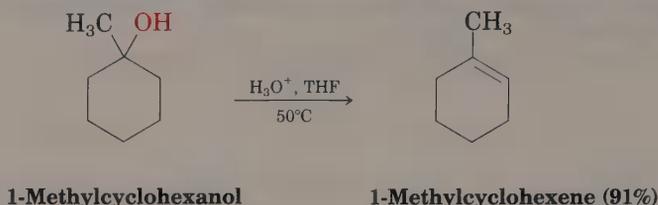
### DEHYDRATION OF ALCOHOLS TO YIELD ALKENES

One of the most important C–O bond reactions of alcohols is **dehydration** to give alkenes. The carbon–oxygen bond is broken, a neighboring C–H bond is broken, and an alkene pi bond is formed:

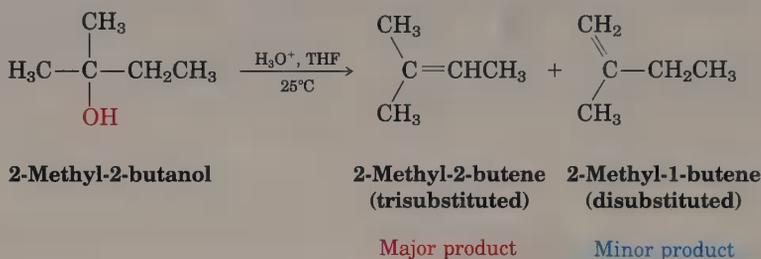


Because of the importance of the reaction, a number of alternative ways have been devised for carrying out dehydrations. One of the more common methods, which works particularly well for tertiary alcohols, is the acid-catalyzed method discussed in Section 7.1. For example, when 1-methyl-

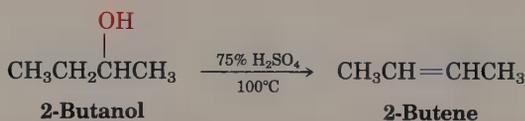
cyclohexanol is warmed with aqueous sulfuric acid in a solvent such as tetrahydrofuran, loss of water occurs and 1-methylcyclohexene is formed:



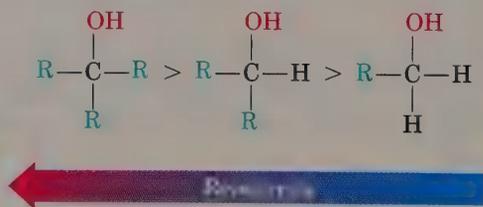
Acid-catalyzed dehydrations normally follow Zaitsev's rule (Section 11.10) and yield the more highly substituted alkene as the major product. Thus, 2-methyl-2-butanol gives primarily 2-methyl-2-butene (trisubstituted double bond) rather than 2-methyl-1-butene (disubstituted double bond):



In normal laboratory practice, only tertiary alcohols are commonly dehydrated with acid. Secondary alcohols can be made to react, but the conditions are severe (75%  $\text{H}_2\text{SO}_4$ ,  $100^\circ\text{C}$ ) and sensitive molecules don't survive.

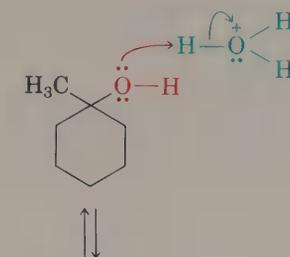


Primary alcohols are even less reactive than secondary ones, and very harsh conditions are necessary to cause dehydration (95%  $\text{H}_2\text{SO}_4$ ,  $150^\circ\text{C}$ ). Thus, the reactivity order for acid-catalyzed dehydrations is

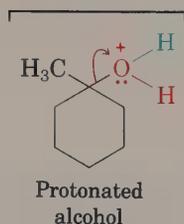


The reasons for the observed reactivity order are best understood by looking at the mechanism of the reaction (Figure 17.8). Acid-catalyzed dehydrations are  $\text{E1}$  reactions (Section 11.14), which occur by a three-step mechanism involving protonation of the alcohol oxygen, loss of water to generate a carbocation intermediate, and final loss of a proton ( $\text{H}^+$ ) from the neighboring carbon atom.

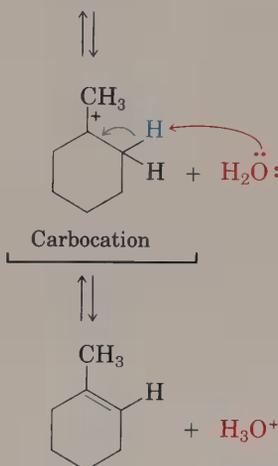
Two electrons from the oxygen atom bond to  $\text{H}^+$ , yielding a protonated alcohol intermediate.



The carbon–oxygen bond breaks, and the two electrons from the bond stay with oxygen, leaving a carbocation intermediate.



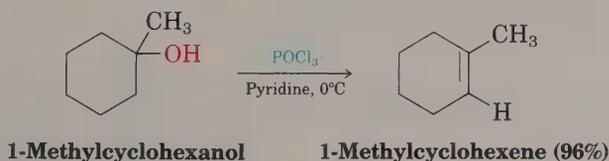
Two electrons from a neighboring carbon–hydrogen bond form the alkene pi bond, and  $\text{H}^+$  (a proton) is eliminated.



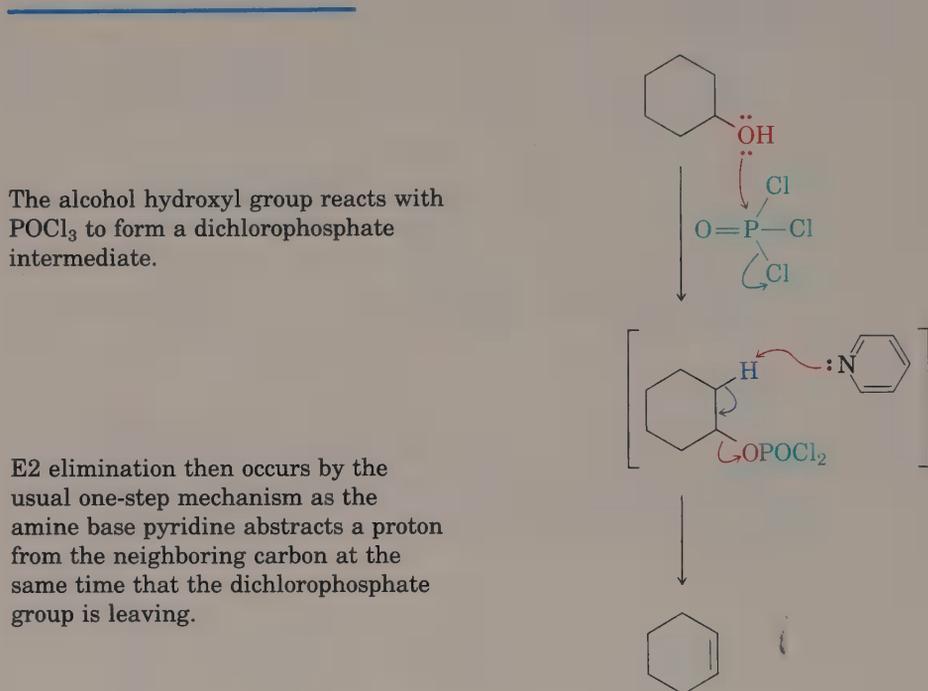
**Figure 17.8** Mechanism of the acid-catalyzed dehydration of alcohols to yield alkenes.

Once the acid-catalyzed dehydration is recognized to be an E1 reaction, the reason why tertiary alcohols react fastest becomes clear. Tertiary substrates *always* react fastest in E1 reactions because they lead to highly stabilized tertiary carbocation intermediates.

To circumvent the necessity for strong acids and allow the dehydration of secondary alcohols in a gentler way, other reagents have been developed that are effective under mild, basic conditions. One such reagent, phosphorus oxychloride ( $\text{POCl}_3$ ), is often able to effect the dehydration of secondary and tertiary alcohols at  $0^\circ\text{C}$  in the basic amine solvent pyridine:



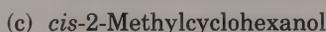
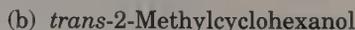
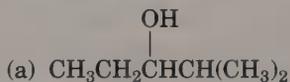
Alcohol dehydrations carried out with  $\text{POCl}_3$  take place by the E2 mechanism shown in Figure 17.9. Since hydroxide ion is a poor leaving group (Section 11.5), direct E2 elimination of water from an alcohol doesn't occur. In the presence of  $\text{POCl}_3$ , however, the hydroxyl group is converted into a dichlorophosphate, which is an excellent leaving group and is readily eliminated to yield an alkene. Pyridine, an organic amine, serves both as reaction solvent and as base to abstract a neighboring proton in the E2 elimination step.



**Figure 17.9** Mechanism of the dehydration of secondary and tertiary alcohols by reaction with  $\text{POCl}_3$  in pyridine. The reaction is an E2 process.

**PROBLEM** .....

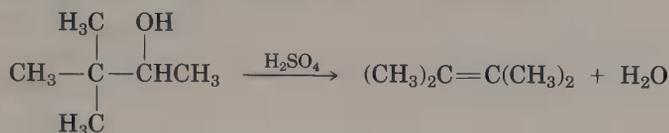
- 17.13** What product(s) would you expect to obtain from dehydration of these alcohols with  $\text{POCl}_3$  in pyridine? Indicate the major product in each case.



**PROBLEM** .....

- 17.14** Good evidence for the intermediacy of carbocations in the acid-catalyzed dehydration of alcohols comes from the observation that rearrangements sometimes occur. Pro-

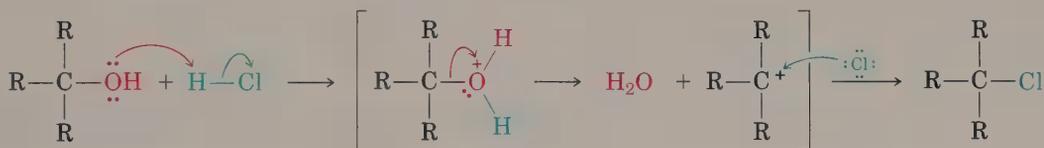
pose a mechanism to account for the formation of 2,3-dimethyl-2-butene from 3,3-dimethyl-2-butanol. [Hint: See Section 6.13.]



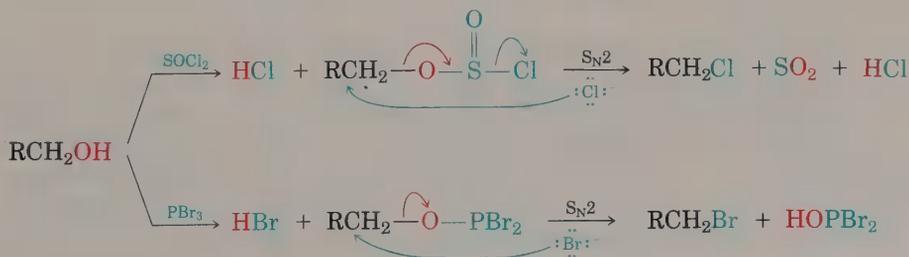
## CONVERSION OF ALCOHOLS INTO ALKYL HALIDES

A second C–O bond reaction of alcohols is their conversion into alkyl halides (Section 10.8). Tertiary alcohols are readily converted into alkyl halides by treatment with either HCl or HBr at 0°C. Primary and secondary alcohols are much more resistant to acid, however, and are best converted into halides by treatment with either SOCl<sub>2</sub> or PBr<sub>3</sub>.

As discussed in Section 11.16, the reaction of a tertiary alcohol with HX takes place by an S<sub>N</sub>1 route. Acid protonates the hydroxyl oxygen atom, water is expelled to generate a carbocation, and the cation reacts with nucleophilic halide ion to give the alkyl halide product:



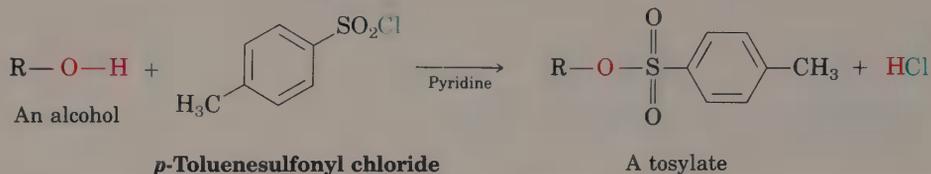
The reactions of primary and secondary alcohols with SOCl<sub>2</sub> and PBr<sub>3</sub> take place by S<sub>N</sub>2 routes. Hydroxide ion itself is too poor a leaving group to be displaced by nucleophiles in S<sub>N</sub>2 reactions, but reaction of an alcohol with SOCl<sub>2</sub> or PBr<sub>3</sub> converts the hydroxyl into a much better leaving group that is readily expelled by back-side nucleophilic attack (Figure 17.10).



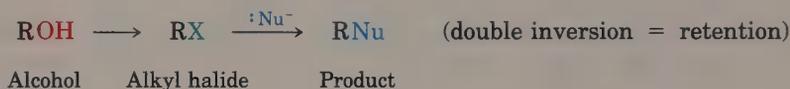
**Figure 17.10** Conversion of a primary alcohol into alkyl halides by S<sub>N</sub>2 reactions with SOCl<sub>2</sub> and PBr<sub>3</sub>.

## CONVERSION OF ALCOHOLS INTO TOSYLATES

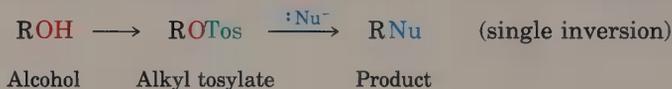
We saw in Section 11.2 that alcohols react with *p*-toluenesulfonyl chloride (tosyl chloride, *p*-TosCl) in pyridine solution to yield alkyl tosylates, ROTos. Only the O–H bond of the alcohol is broken in this reaction; the C–O bond remains intact, and no change of configuration occurs if the alcohol is chiral. The resultant alkyl tosylates behave much like alkyl halides in their chemistry, undergoing  $S_N1$  and  $S_N2$  substitution reactions with ease.



One of the most important reasons for using tosylates instead of halides is stereochemical. In a sense, the following two sequences of reactions are stereochemically complementary.



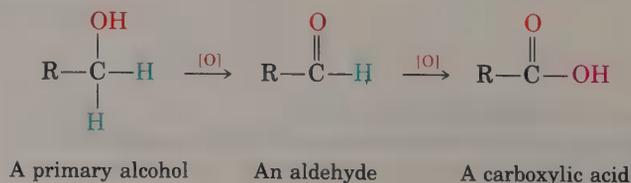
and

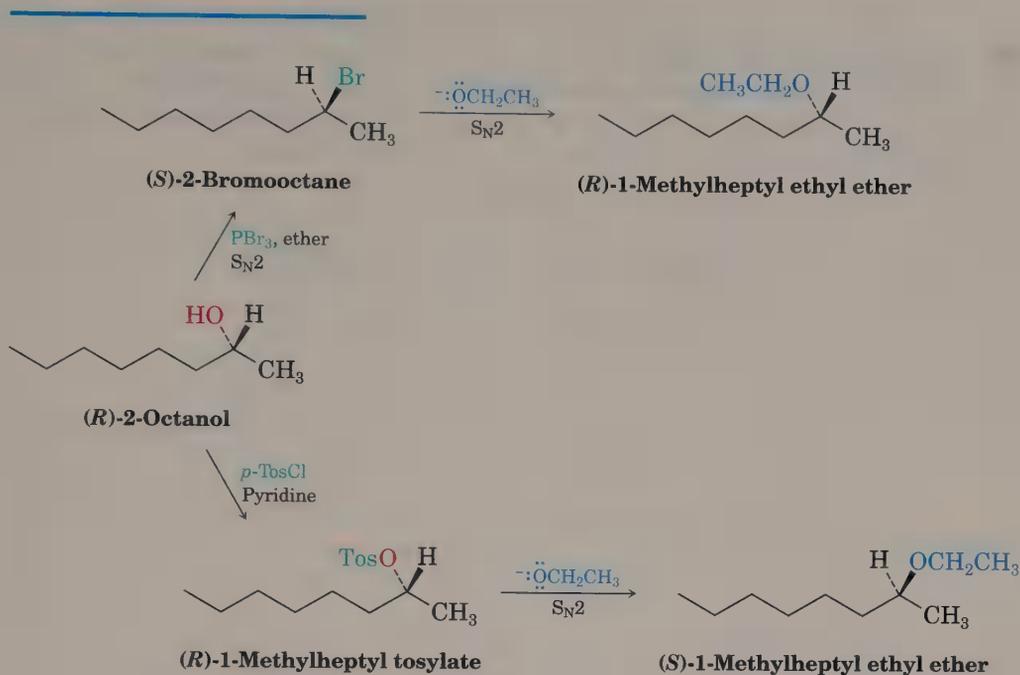


The  $S_N2$  reaction via the halide proceeds with *two* Walden inversions—one to make the halide from the alcohol and one to substitute the halide—and yields a product with the same absolute stereochemistry as the starting material. The  $S_N2$  reaction via the tosylate proceeds with only *one* Walden inversion, however, and yields a product of opposite absolute stereochemistry from the starting material. Figure 17.11 gives a series of reactions on optically active 2-octanol that illustrates these stereochemical relationships.

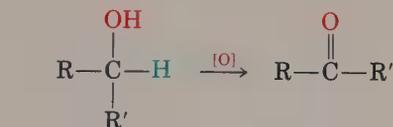
## 17.9 Oxidation of Alcohols

The most important reaction of alcohols is their oxidation to yield carbonyl compounds. Primary alcohols yield aldehydes or carboxylic acids; secondary alcohols yield ketones; and tertiary alcohols don't react with most oxidizing agents except under the most vigorous conditions.



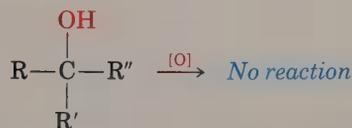


**Figure 17.11** Stereochemical consequences of some  $\text{S}_\text{N}2$  reactions on derivatives of  $(R)$ -2-octanol.



A secondary alcohol

A ketone

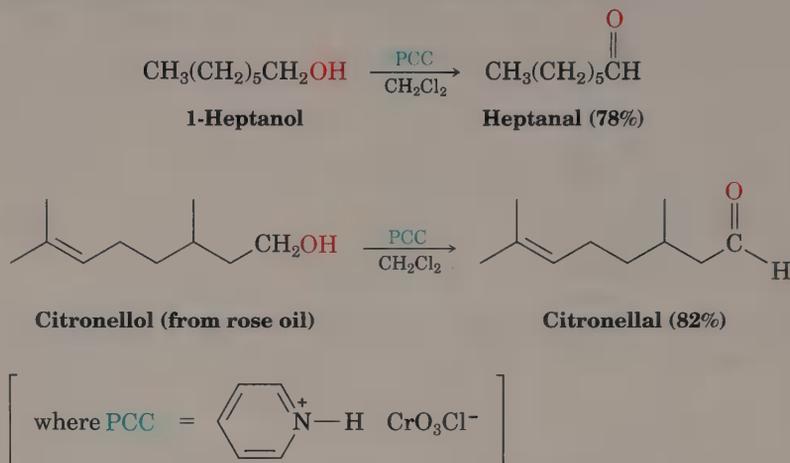


A tertiary alcohol

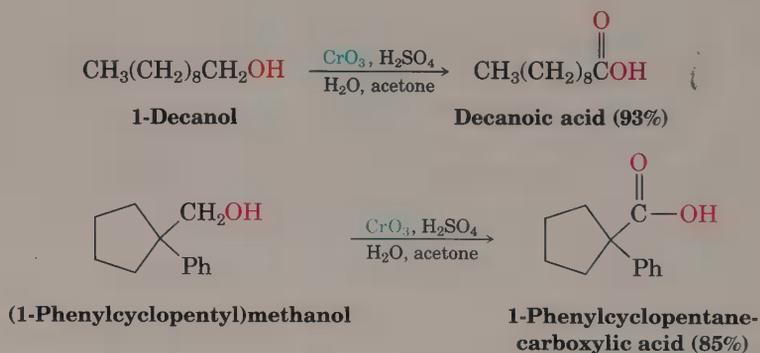
where  $[\text{O}]$  is an oxidizing agent

Oxidation of primary and secondary alcohols can be accomplished by a large number of reagents, including  $\text{KMnO}_4$ ,  $\text{CrO}_3$ , and  $\text{Na}_2\text{Cr}_2\text{O}_7$ . Which reagent is used in a specific case depends on such factors as cost, convenience, reaction yield, and alcohol sensitivity. For example, the large-scale oxidation of a simple, inexpensive alcohol like cyclohexanol would probably best be done with a cheap oxidant such as potassium permanganate. On the other hand, the small-scale oxidation of a delicate and expensive polyfunctional alcohol would best be done with a mild and high-yielding reagent, regardless of cost.

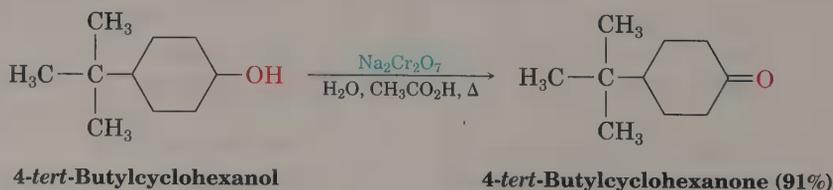
Primary alcohols are oxidized either to aldehydes or to carboxylic acids, depending on the reagents chosen and on the conditions used. Probably the best method for preparing aldehydes from primary alcohols on a laboratory scale (as opposed to an industrial scale) is by use of pyridinium chlorochromate (PCC,  $C_5H_6NCrO_3Cl$ ) in dichloromethane solvent.



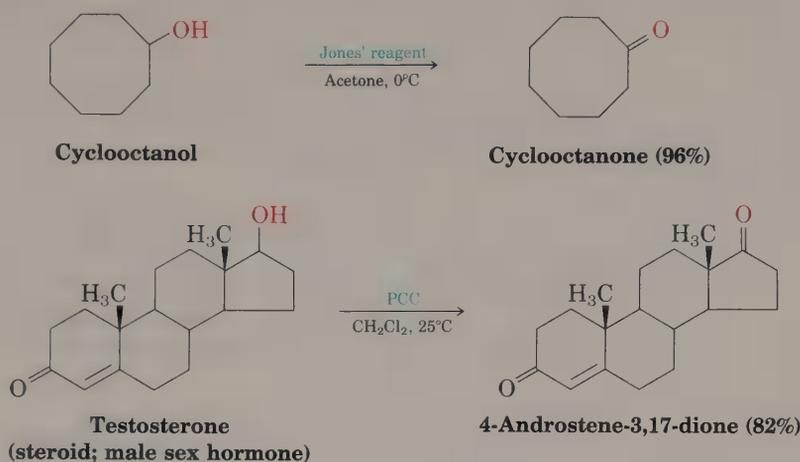
Most other oxidizing agents, such as chromium trioxide ( $CrO_3$ ) in aqueous sulfuric acid (*Jones' reagent*), oxidize primary alcohols to carboxylic acids. Aldehydes are involved as intermediates in the Jones oxidation but can't usually be isolated because they are further oxidized too readily.



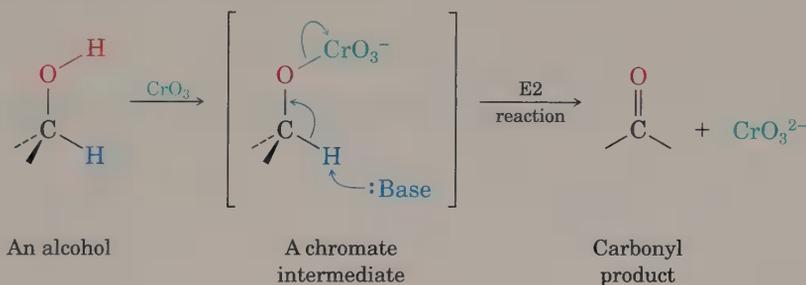
Secondary alcohols are oxidized easily and in high yields to give ketones. For large-scale oxidations, an inexpensive reagent such as sodium dichromate in aqueous acetic acid is used.



For more sensitive alcohols, pyridinium chlorochromate or Jones' reagent is often used, since these reactions are milder and occur at lower temperatures.



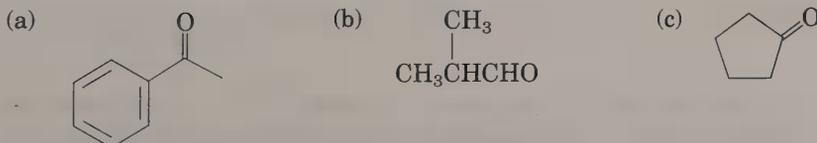
All of these oxidations occur by a pathway closely related to the E2 reaction. The first step involves reaction between the alcohol and a chromium(VI) reagent to form an intermediate chromate. Bimolecular elimination then yields the carbonyl product.



Although we usually think of the E2 reaction as a means of generating carbon-carbon double bonds by dehydrohalogenation of alkyl halides, it's also useful for preparing carbon-oxygen double bonds. This is just one more example of how the same few fundamental mechanistic types keep reappearing in different variations.

PROBLEM .....

17.15 Which alcohols would give these products on oxidation?

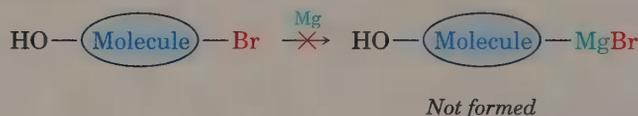


## PROBLEM.....

- 17.16 What products would you expect from oxidation of these compounds with Jones' reagent? With pyridinium chlorochromate?
- (a) 1-Hexanol                      (b) 2-Hexanol                      (c) Hexanal

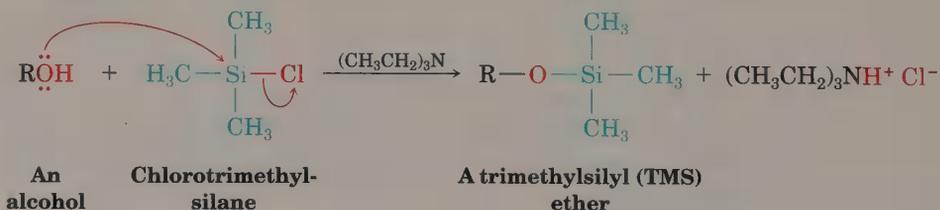
## 17.10 Protection of Alcohols

It often happens, particularly during the synthesis of complex molecules, that one functional group in a molecule interferes with an intended reaction on a second functional group elsewhere in the same molecule. For example, we saw earlier in this chapter that a Grignard reagent can't be prepared from a halo alcohol because the carbon–magnesium bond is not compatible with the presence of an acidic hydroxyl group in the same molecule.

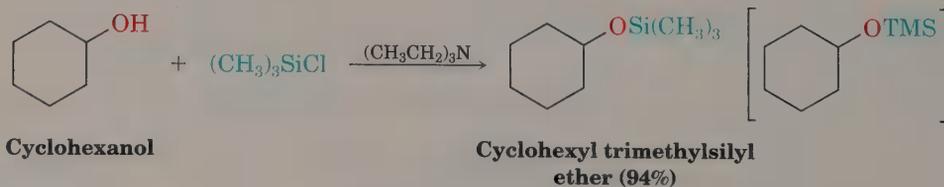


When this kind of incompatibility arises, it's sometimes possible to circumvent the problem by *protecting* the interfering functional group. Protection involves three steps: (1) introducing a **protecting group** to form an inert derivative, (2) carrying out the desired reaction, and (3) removal of the protecting group.

One of the most common methods of alcohol protection involves base-catalyzed reaction with chlorotrimethylsilane to yield a trimethylsilyl (TMS) ether. The reaction is an  $S_N2$  reaction of the alcohol anion with the chlorosilane:

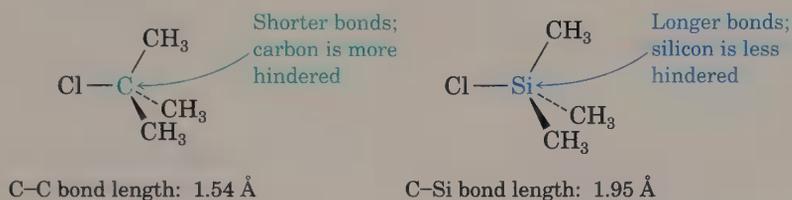


For example:

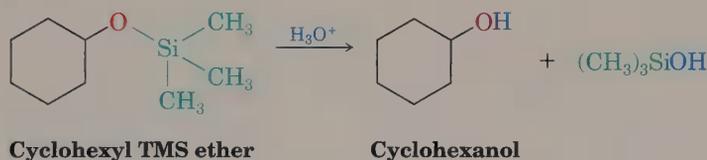


As indicated by the preceding example, TMS ether formation is carried out in the presence of a base, usually triethylamine, to help form the alkoxide anion from the alcohol and scavenge the HCl by-product from the reaction.

The actual ether-forming step is an  $S_N2$  back-side attack of the alcohol on the silicon atom, with concurrent loss of a leaving chloride anion. Unlike most  $S_N2$  reactions, however, which are strongly affected by steric hindrance, this one takes place at a *tertiary* center—a trialkyl substituted silicon atom. The reaction occurs because silicon, a second-row atom, is larger than carbon and thus forms longer bonds. As a result, the three methyl substituents attached to silicon are farther away and offer less steric hindrance to  $S_N2$  attack than they do in the analogous *tert*-butyl chloride.



Like most other ethers that we'll study in the next chapter, TMS ethers are relatively unreactive. They have no acidic protons and are therefore protected against reaction with oxidizing agents, reducing agents, and Grignard reagents. They can, however, be cleaved by reaction with aqueous acid or with fluoride ion to regenerate the alcohol. The cleavage reaction is simply another  $S_N2$  attack on silicon.



To complete the earlier example, it's possible to use a halo alcohol in a Grignard reaction by employing a three-step sequence: (1) protection, (2) Grignard formation and reaction, (3) deprotection. For example, we can add 3-bromo-1-propanol to acetaldehyde by the route shown in Figure 17.12.

#### PROBLEM .....

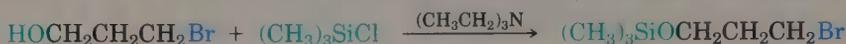
- 17.17 TMS ethers can be removed by treatment with fluoride ion as well as by acid-catalyzed hydrolysis. Propose a mechanism for the reaction of cyclohexyl TMS ether with LiF. Fluorotrimethylsilane is a product.
- .....

## 17.11 Spectroscopic Analysis of Alcohols

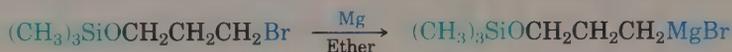
### INFRARED SPECTROSCOPY

Alcohols show a characteristic O-H stretching absorption at 3300–3600  $cm^{-1}$  in the infrared spectrum that simplifies their spectroscopic identification. The exact position of the absorption band depends on the extent of

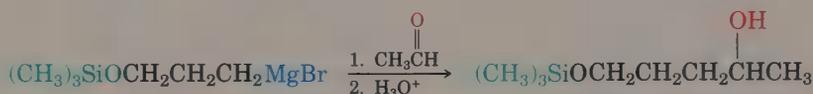
Step 1: Protect alcohol



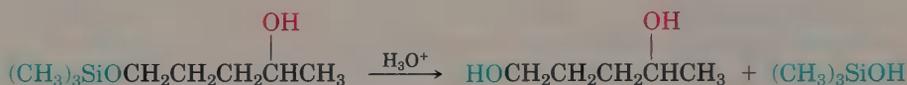
Step 2a: Form Grignard reagent



Step 2b: Do Grignard reaction

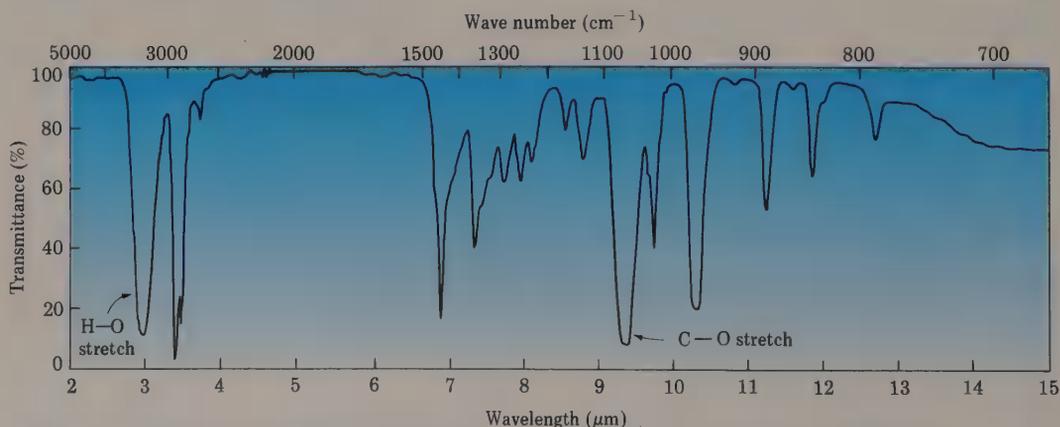


Step 3: Remove protecting group



**Figure 17.12** Use of a TMS-protected alcohol during a Grignard reaction.

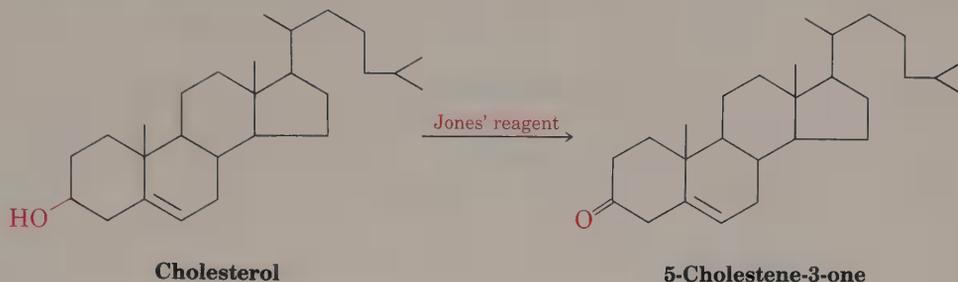
hydrogen bonding in the sample. Unassociated alcohols show a fairly sharp absorption near  $3600\text{ cm}^{-1}$ , whereas hydrogen-bonded alcohols show a broader absorption in the  $3300\text{--}3400\text{ cm}^{-1}$  range. The hydrogen-bonded hydroxyl absorption is easily seen at  $3350\text{ cm}^{-1}$  in the infrared spectrum of cyclohexanol (Figure 17.13). Alcohols also show a strong C–O stretching absorption near  $1050\text{ cm}^{-1}$ .



**Figure 17.13** Infrared spectrum of cyclohexanol.

## PROBLEM.....

- 17.18 Let's assume that you needed to prepare 5-cholestene-3-one from cholesterol by Jones oxidation. How could you use infrared spectroscopy to tell if the reaction was successful? What differences would you look for in the infrared spectra of starting material and product?



## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Carbon atoms bonded to electron-withdrawing  $-OH$  groups are deshielded and absorb at a lower field than do normal alkane carbon atoms in the  $^{13}C$  NMR spectrum. Most alcohol carbon absorptions fall in the range 50–80  $\delta$ , as the spectral data in Figure 17.14 illustrate for cyclohexanol.

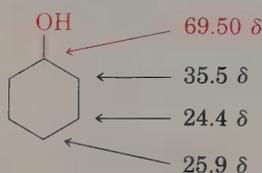
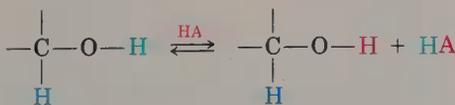


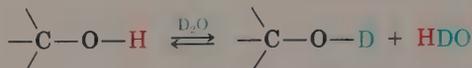
Figure 17.14 Carbon-13 NMR data for cyclohexanol.

Alcohols show characteristic absorptions in the  $^1H$  NMR spectrum also. Protons on the oxygen-bearing carbon atom are deshielded by the electron-withdrawing effect of the nearby oxygen, and their absorptions occur in the range 3.5–4.5  $\delta$ . Surprisingly, splitting is not usually observed between the  $O-H$  proton and the neighboring protons on carbon. Most samples contain small amounts of acidic impurities that catalyze an exchange of the hydroxyl proton on a time scale so rapid that the effects of spin–spin splitting are removed and no coupling is observed.

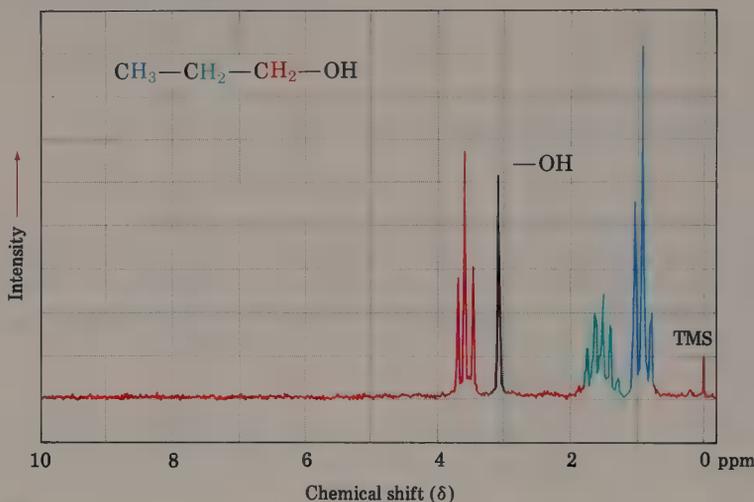


No NMR coupling observed

It's often possible to take advantage of this rapid proton exchange to identify the position of the O–H absorption. If a small amount of deuterated water,  $D_2O$ , is added to the NMR sample tube, the O–H proton is rapidly exchanged for deuterium, and the hydroxyl absorption disappears from the spectrum.



Normal spin–spin splitting is observed between protons on the oxygen-bearing carbon and other neighbors. For example, the signal of the two  $-\text{CH}_2\text{O}-$  protons in 1-propanol is split into a triplet by coupling with the neighboring  $-\text{CH}_2-$  protons (Figure 17.15).



**Figure 17.15** Proton NMR spectrum of 1-propanol. The protons on the oxygen-bearing carbon are split into a triplet at 3.6  $\delta$ .

#### PROBLEM

- 17.19** When the  $^1\text{H}$  NMR spectra of alcohols are run in dimethyl sulfoxide (DMSO) solvent, exchange of the hydroxyl proton is slow, and spin–spin splitting is observed between the  $-\text{O}-\text{H}$  proton and  $\text{C}-\text{H}$  protons on the adjacent carbon. What spin multiplicities would you expect for the hydroxyl protons in these alcohols?
- (a) 2-Methyl-2-propanol      (b) Cyclohexanol      (c) Ethanol  
 (d) 2-Propanol      (e) Cholesterol      (f) 1-Methylcyclohexanol

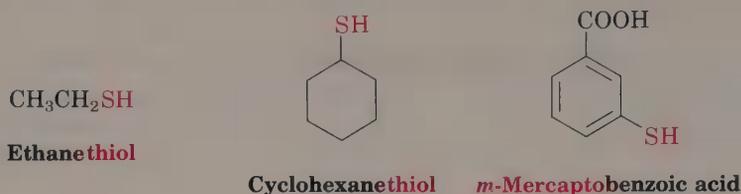
## MASS SPECTROSCOPY

Alcohols undergo fragmentation in the mass spectrometer by two characteristic pathways, *alpha cleavage* and *dehydration*. In the alpha-cleavage



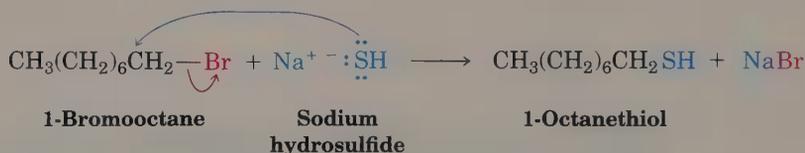
## 17.12 Thiols

**Thiols, R-SH**, are sulfur analogs of alcohols. Thiols are named by the same system used for alcohols, with the suffix *-thiol* used in place of *-ol*. The -SH group itself is referred to as a **mercapto** group.

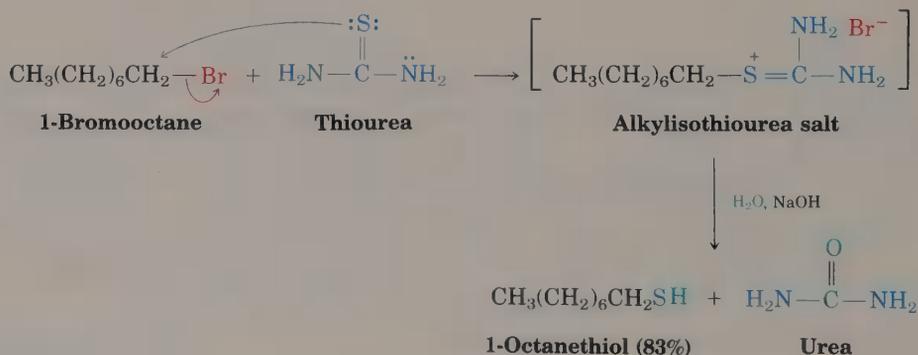


The most obvious physical characteristic of thiols is their appalling odor. For example, skunk scent is caused primarily by the simple thiols, 3-methyl-1-butanethiol and 2-butene-1-thiol. Small amounts of low-molecular-weight thiols are added to natural gas to serve as an easily detectable warning in case of leaks.

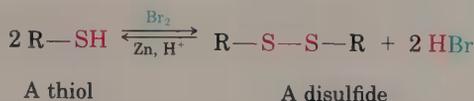
Thiols are usually prepared from the corresponding alkyl halides by  $\text{S}_{\text{N}}2$  displacement with a sulfur nucleophile such as hydrosulfide anion,  $\text{SH}^-$ .



Yields are often poor in this reaction unless an excess of hydrosulfide anion is used, because the product thiol can undergo further reaction with alkyl halide, yielding a symmetrical sulfide, R-S-R, as a by-product. For this reason, thiourea,  $(\text{NH}_2)_2\text{C}=\text{S}$ , is often used as the nucleophile in the preparation of thiols from alkyl halides. The reaction occurs by displacement of the halide ion to yield an intermediate alkyliothiurea salt, followed by hydrolysis with aqueous base.



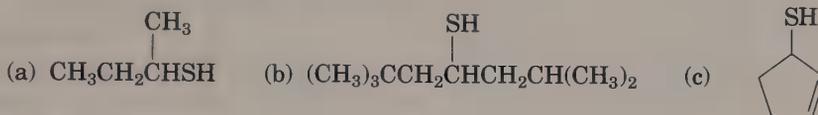
Thiols can be oxidized by mild reagents such as bromine or iodine to yield **disulfides, RSSR**. The reaction is easily reversed, and disulfides can be reduced back to thiols by treatment with zinc and acid.



We'll see later that the thiol–disulfide interconversion is extremely important in biochemistry, where disulfide “bridges” form the cross-links between protein chains that help stabilize the three-dimensional conformations of proteins.

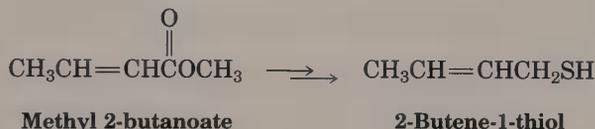
PROBLEM .....

17.20 Name the following compounds by IUPAC rules:



PROBLEM .....

17.21 2-Butene-1-thiol is one component of skunk spray. (a) How would you synthesize this substance from methyl 2-butenate? (b) From 1,3-butadiene?



## 17.13 Summary and Key Words

**Alcohols** are among the most versatile of all organic compounds. They occur widely in nature, are important industrially, and have an unusually rich chemistry. Alcohols can be prepared in a number of different ways, including osmium tetroxide **hydroxylation** of alkenes, **hydroboration–oxidation** of alkenes, and **oxymercuration**–sodium borohydride reduction of alkenes.

The most important methods of alcohol synthesis involve carbonyl compounds. Aldehydes, ketones, esters, and carboxylic acids can be **reduced** by reaction with either sodium borohydride or lithium aluminum hydride. Aldehydes, esters, and carboxylic acids yield primary alcohols on reduction; ketones yield secondary alcohols. **Sodium borohydride** is safe to use and relatively unreactive. It reduces aldehydes and ketones rapidly, reduces esters slowly, and does not reduce carboxylic acids. **Lithium aluminum hydride** is a powerful and dangerous reagent that must be handled with great care. It rapidly reduces all types of carbonyl groups, including carboxylic acids.

The reaction of **Grignard reagents** with carbonyl compounds is another important method for preparing alcohols. Grignard addition to formaldehyde

yields a primary alcohol, addition to an aldehyde yields a secondary alcohol, and addition to a ketone or an ester yields a tertiary alcohol. Carboxylic acids don't give Grignard addition products. The Grignard synthesis of alcohols is limited by the fact that Grignard reagents can't be prepared from alkyl halides that contain reactive functional groups in the same molecule. This problem can sometimes be avoided by **protecting** the interfering functional group. Alcohols are often protected by formation of **trimethylsilyl (TMS) ethers**.

Alcohols undergo a great many different reactions. They can be dehydrated by treatment with  $\text{POCl}_3$  and can be transformed into alkyl halides by treatment with  $\text{PBr}_3$  or  $\text{SOCl}_2$ . Furthermore, alcohols are weakly acidic ( $\text{p}K_a \approx 16\text{--}18$ ). They react with strong bases and with alkali metals to form **alkoxide anions**, which are much used in organic synthesis.

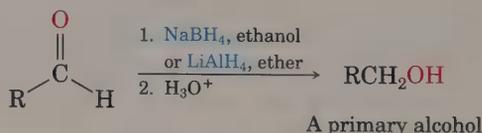
The most important reaction of alcohols is their **oxidation** to carbonyl compounds. Primary alcohols yield either aldehydes or carboxylic acids, depending on the conditions used; secondary alcohols yield ketones; and tertiary alcohols aren't normally oxidized. **Pyridinium chlorochromate (PCC)** in dichloromethane is often used for oxidizing primary alcohols to aldehydes and secondary alcohols to ketones. The **Jones reagent**, a solution of  $\text{CrO}_3$  in aqueous sulfuric acid, is much used for oxidizing primary alcohols to carboxylic acids and secondary alcohols to ketones.

Alcohols show characteristic absorptions in both infrared and nuclear magnetic resonance spectra, making their spectroscopic analysis relatively straightforward.

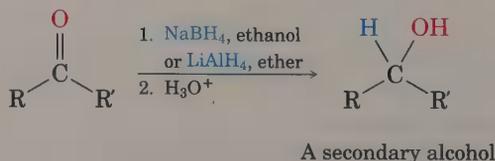
**Thiols,  $\text{RSH}$** , the sulfur analogs of alcohols, are usually prepared by  $\text{S}_\text{N}2$  reaction of an alkyl halide with thiourea. Mild oxidation of a thiol yields a **disulfide,  $\text{RSSR}$** , and mild reduction of a disulfide gives back the thiol.

## 17.14 Summary of Reactions

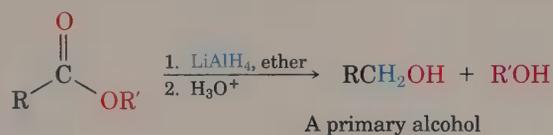
1. Synthesis of alcohols
  - a. Reduction of carbonyl compounds (Section 17.6)
    - (1) Aldehydes



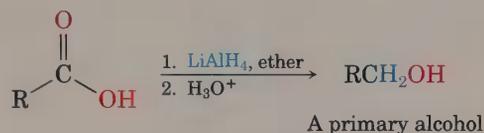
- (2) Ketones



## (3) Esters

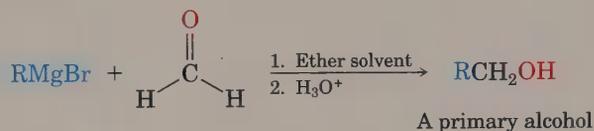


## (4) Carboxylic acids

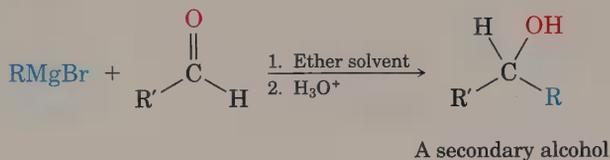


## b. Grignard addition to carbonyl compounds (Section 17.7)

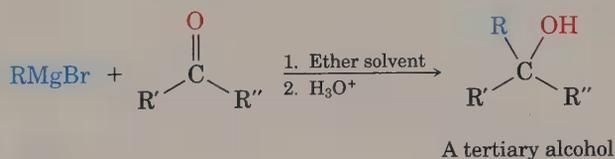
## (1) Formaldehyde



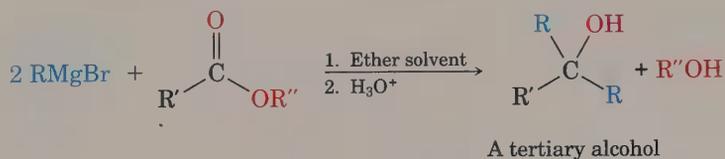
## (2) Aldehydes



## (3) Ketones

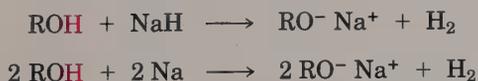


## (4) Esters



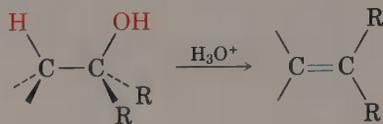
## 2. Reactions of alcohols

## a. Acidity (Section 17.4)



## b. Dehydration (Section 17.8)

## (1) Tertiary alcohols

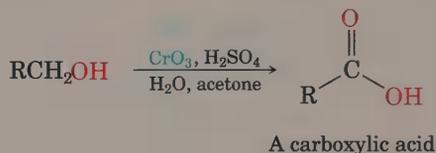
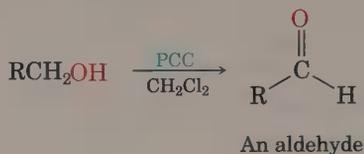


## (2) Secondary and tertiary alcohols

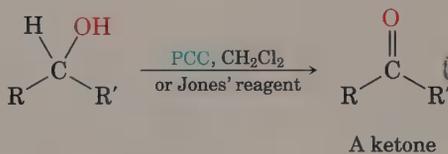


## c. Oxidation (Section 17.9)

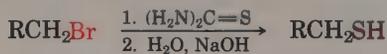
## (1) Primary alcohol



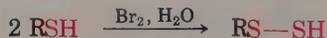
## (2) Secondary alcohol



## 3. Synthesis of thiols (Section 17.12)

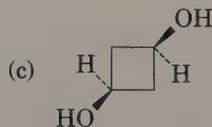
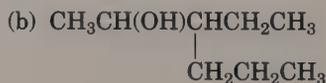
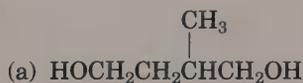


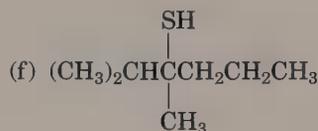
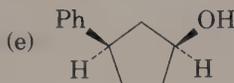
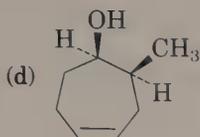
## 4. Oxidation of thiols to disulfides (Section 17.12)



## ADDITIONAL PROBLEMS

17.22 Name the following compounds according to the IUPAC system:





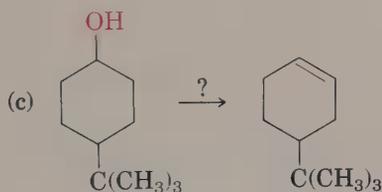
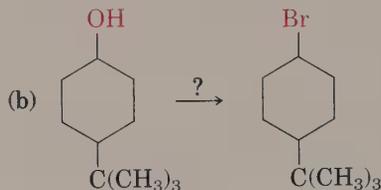
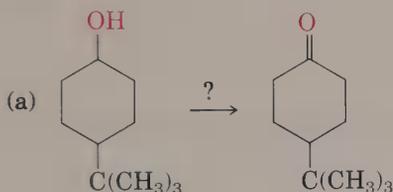
17.23 Draw and name the eight isomeric alcohols with formula  $\text{C}_5\text{H}_{12}\text{O}$ .

17.24 Which of the eight alcohols you identified in Problem 17.23 react with Jones' reagent ( $\text{CrO}_3$ ,  $\text{H}_2\text{O}$ ,  $\text{H}_2\text{SO}_4$ )? Show the products you would expect from each reaction.

17.25 How would you prepare the following compounds from 2-phenylethanol?

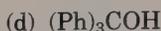
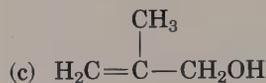
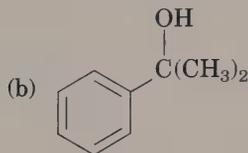
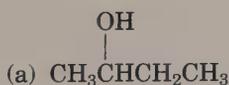
- |  |  |
|--|--|
| (a) Styrene  | (b) Phenylacetaldehyde ( $\text{C}_6\text{H}_5\text{CH}_2\text{CHO}$ ) |
| (c) Phenylacetic acid ( $\text{C}_6\text{H}_5\text{CH}_2\text{CO}_2\text{H}$ ) | (d) Benzoic acid   |
| (e) Ethylbenzene   | (f) Benzaldehyde   |
| (g) 1-Phenylethanol  | (h) 1-Bromo-2-phenylethane   |

17.26 How would you carry out these transformations?



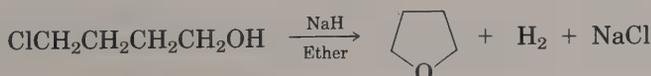
- |  |                   |   |
|--|-------------------|---|
| (d) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ | $\xrightarrow{?}$ | $\text{CH}_3\text{CH}_2\text{CHO}$                        |
| (e) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ | $\xrightarrow{?}$ | $\text{CH}_3\text{CH}_2\text{CO}_2\text{H}$               |
| (f) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ | $\xrightarrow{?}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{OTos}$            |
| (g) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ | $\xrightarrow{?}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}$              |
| (h) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ | $\xrightarrow{?}$ | $\text{CH}_3\text{CH}_2\text{CH}_2\text{O}^- \text{Na}^+$ |

17.27 What Grignard reagent and what carbonyl compound might you start with to prepare these alcohols?

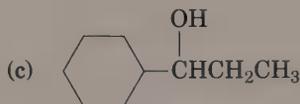
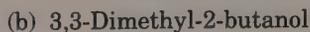
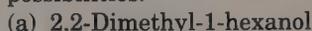


17.28 Assume that you have been given a sample of (*S*)-2-octanol. How could you prepare (*R*)-2-chlorooctane? How could you prepare (*R*)-2-octanol?

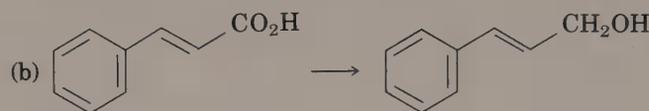
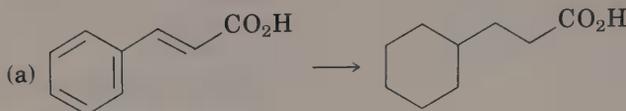
17.29 When 4-chloro-1-butanol is treated with a strong base such as sodium hydride,  $\text{NaH}$ , tetrahydrofuran is produced. Suggest a mechanism for this reaction.



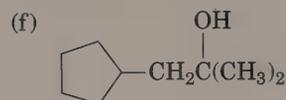
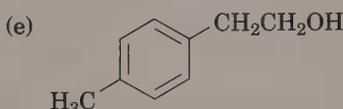
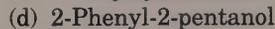
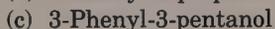
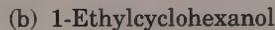
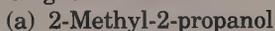
17.30 What carbonyl compounds would you reduce to prepare these alcohols? List all possibilities.



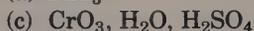
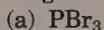
17.31 How would you carry out these transformations?



17.32 What carbonyl compounds might you start with to prepare these compounds by Grignard reaction? List all possibilities.



17.33 What products would you expect to obtain from reaction of 1-pentanol with these reagents?

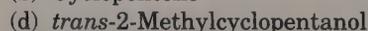
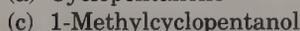
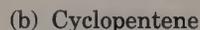
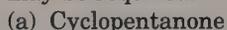


17.34 Acid-catalyzed dehydration of 2,2-dimethylcyclohexanol yields a mixture of 1,2-dimethylcyclohexene and isopropylidenecyclopentane. Propose a mechanism to account for the formation of both products.

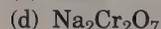
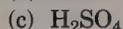
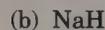
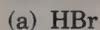


Isopropylidenecyclopentane

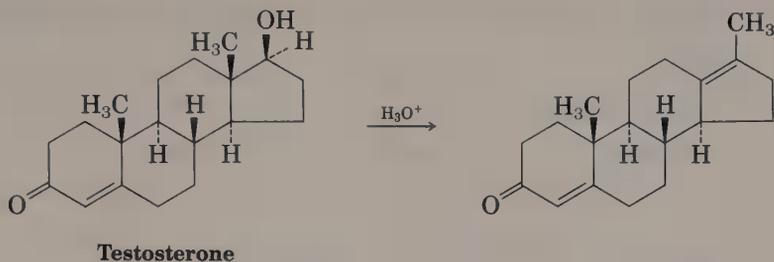
17.35 How would you prepare these substances from cyclopentanol? More than one step may be required.



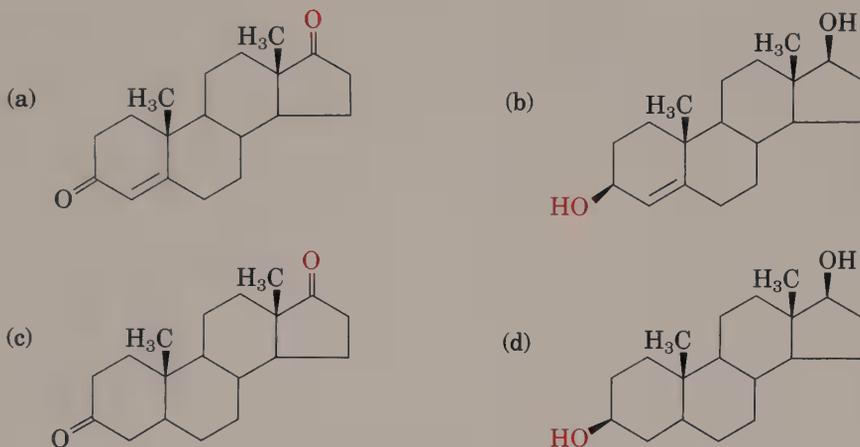
17.36 What products would you expect to obtain from reaction of 1-methylcyclohexanol with these reagents?



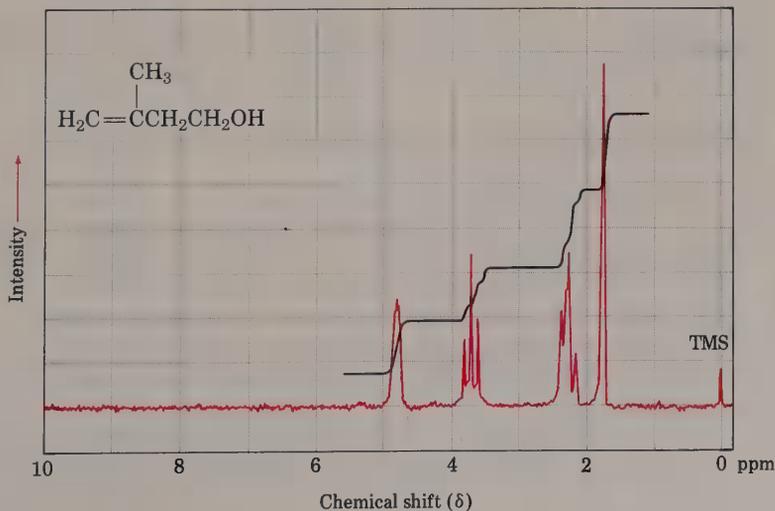
17.37 Testosterone is one of the most important male steroid hormones. When testosterone is dehydrated by treatment with acid, rearrangement occurs to yield the product indicated. Propose a mechanism to account for this reaction. [Hint: See Section 6.13.]



- 17.38 Starting from testosterone (Problem 17.37), how would you prepare the following substances?



- 17.39 Compound A,  $\text{C}_{10}\text{H}_{18}\text{O}$ , undergoes reaction with dilute  $\text{H}_2\text{SO}_4$  at  $25^\circ\text{C}$  to yield a mixture of two alkenes,  $\text{C}_{10}\text{H}_{16}$ . The major alkene product, B, gives only cyclopentanone after ozone treatment followed by reduction with zinc in acetic acid. Formulate the reactions involved, and identify A and B.
- 17.40 Dehydration of *trans*-2-methylcyclopentanol with  $\text{POCl}_3$  in pyridine yields predominantly 3-methylcyclopentene. What is the stereochemistry of this dehydration? Can you suggest a reason for formation of the observed product? (Make molecular models!)
- 17.41 The  $^1\text{H}$  NMR spectrum shown is that of 3-methyl-3-buten-1-ol. Assign all of the observed resonance peaks to specific protons, and account for the splitting patterns.

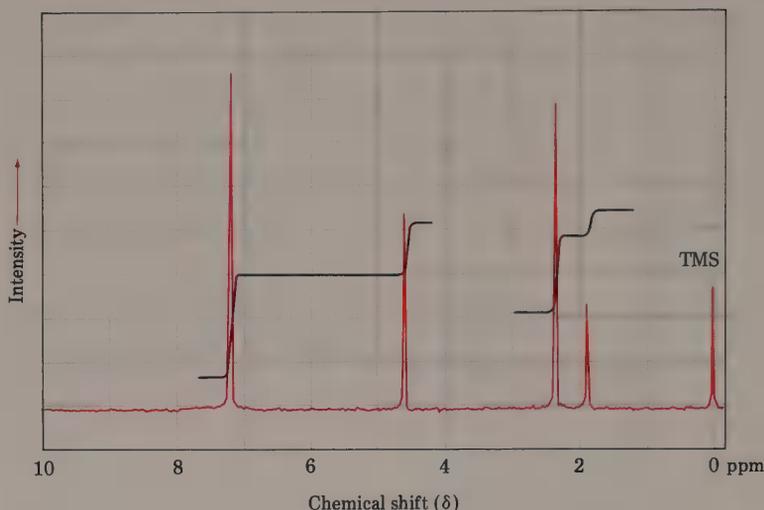


- 17.42 Propose a structure consistent with the following spectral data for a compound of formula  $C_8H_{18}O_2$ :

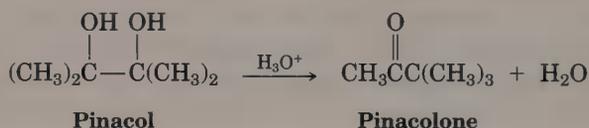
Infrared:  $3350\text{ cm}^{-1}$

$^1\text{H NMR}$ :  $1.24\ \delta$  (12-proton singlet);  $1.56\ \delta$  (4-proton singlet);  
 $1.95\ \delta$  (2-proton singlet)

- 17.43 The  $^1\text{H NMR}$  spectrum shown is of an alcohol of formula  $C_8H_{10}O$ . Propose a structure that is consistent with the observed spectrum.



- 17.44 2,3-Dimethyl-2,3-butanediol has the common name *pinacol*. On heating with aqueous acid, pinacol rearranges to *pinacolone*, 3,3-dimethyl-2-butanone. Can you suggest a mechanism for this reaction? [*Hint*: See Section 6.13.]



- 17.45 Compound A,  $C_5H_{10}O$ , is one of the basic building blocks of nature. All steroids and many other naturally occurring compounds are built up from compound A. Spectroscopic analysis of A yields the following information:

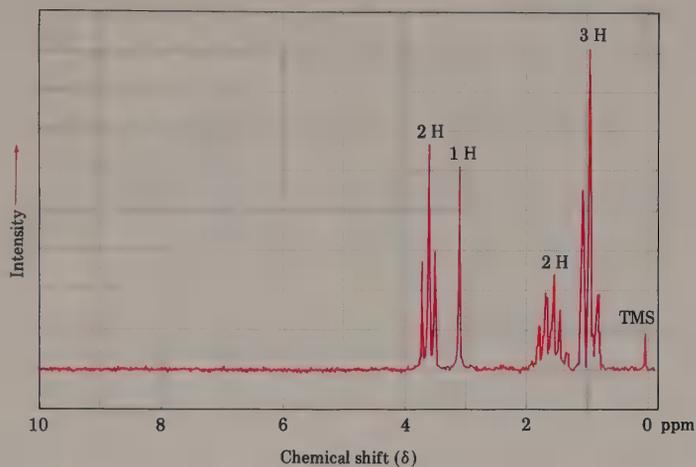
Infrared:  $3400\text{ cm}^{-1}$ ;  $1640\text{ cm}^{-1}$

$^1\text{H NMR}$ :  $1.63\ \delta$  (3-proton singlet);  $1.70\ \delta$  (3-proton singlet);  
 $3.83\ \delta$  (1-proton broad singlet);  $4.15\ \delta$  (2-proton doublet,  $J = 7\text{ Hz}$ );  
 $5.70\ \delta$  (1-proton triplet,  $J = 7\text{ Hz}$ )

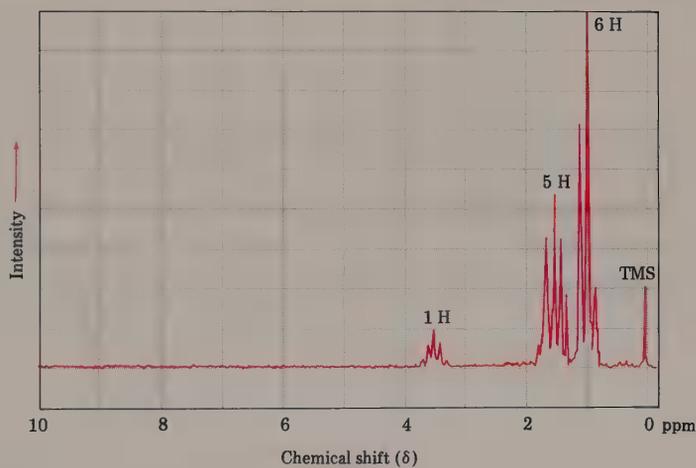
- How many double bonds and/or rings does A have?
- From the infrared spectrum, what is the nature of the oxygen-containing functional group?
- What kinds of protons are responsible for the NMR absorptions listed?
- Propose a structure for A consistent with the  $^1\text{H NMR}$  data.

17.46 Propose structures for alcohols that have the following  $^1\text{H}$  NMR spectra:

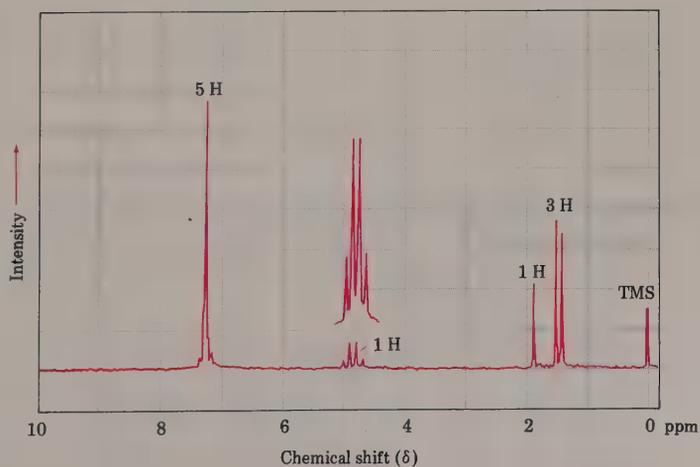
(a)  $\text{C}_3\text{H}_8\text{O}$



(b)  $\text{C}_5\text{H}_{12}\text{O}$

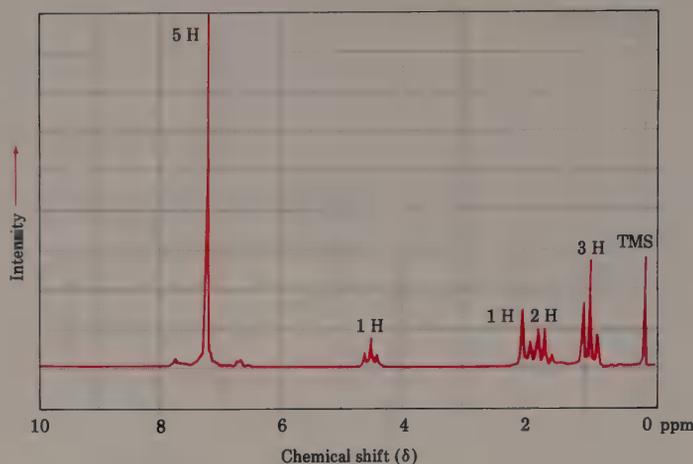


(c)  $\text{C}_8\text{H}_{10}\text{O}$

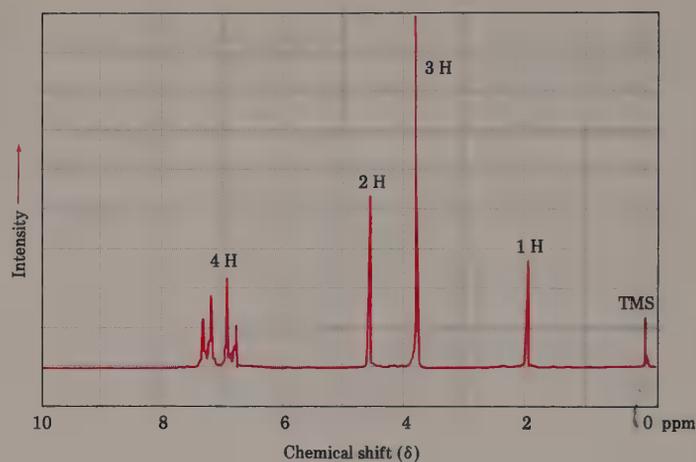


17.47 Propose structures for alcohols that have the following  $^1\text{H}$  NMR spectra:

(a)  $\text{C}_9\text{H}_{12}\text{O}$



(b)  $\text{C}_8\text{H}_{10}\text{O}_2$



- 17.48 As a general rule, axial alcohols oxidize somewhat faster than equatorial alcohols. Which would you expect to oxidize faster, *cis*-4-*tert*-butylcyclohexanol or *trans*-4-*tert*-butylcyclohexanol? Draw the more stable chair conformation of each molecule.
- 17.49 Propose a synthesis of bicyclohexylidene, starting from cyclohexanone as the only source of carbon.



Bicyclohexylidene

- 17.50 Since all hamsters look pretty much alike, attraction between sexes is controlled by chemical secretions. The sex attractant exuded by the female hamster has the following spectral properties:

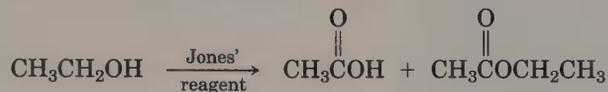
Mass spectrum:  $\text{M}^{+\cdot} = 94$

Infrared: No characteristic functional-group absorptions above  $1500\text{ cm}^{-1}$

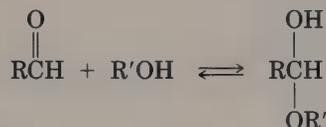
$^1\text{H}$  NMR:  $2.10\ \delta$  (singlet); no other absorptions

Propose a structure for the hamster sex attractant.

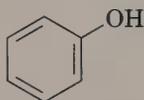
- 17.51 One problem often encountered in the oxidation of primary alcohols to acids is that esters are sometimes produced as by-products. For example, oxidation of ethanol yields acetic acid and ethyl acetate:



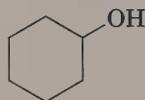
Propose a mechanism to account for the formation of ethyl acetate. Take into account the reversible reaction between aldehydes and alcohols:



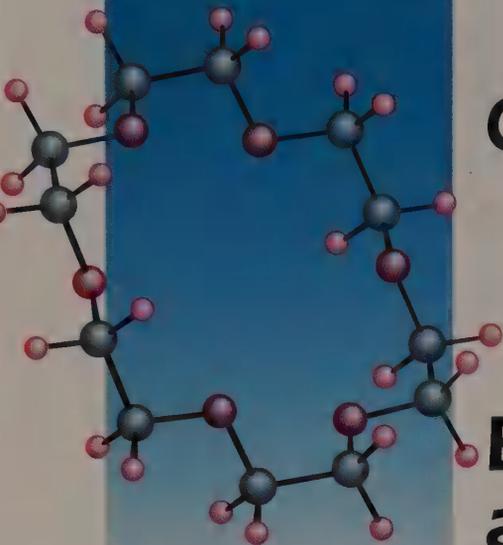
- 17.52 You know the effect of resonance on stability. Use this knowledge to explain why phenol ( $\text{p}K_{\text{a}} = 4.6$ ) is a stronger acid than cyclohexanol ( $\text{p}K_{\text{a}} = 17$ ).



Phenol



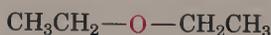
Cyclohexanol



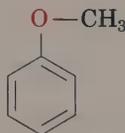
# Chapter 18

## Ethers, Epoxides, and Sulfides

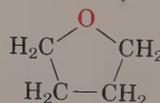
An **ether** is a substance that has two organic residues bonded to the same oxygen atom,  $R-O-R'$ . The organic residues may be alkyl, aryl, or vinylic, and the oxygen atom can be part of either an open chain or a ring. Perhaps the most well known ether is diethyl ether, a familiar substance that has been used medicinally as an anesthetic and is much used industrially as a solvent. Other useful ethers include anisole, a pleasant-smelling aromatic ether used in perfumery, and tetrahydrofuran (THF), a cyclic ether often used as a solvent.



Diethyl ether



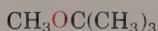
Anisole  
(methyl phenyl ether)



Tetrahydrofuran  
(a cyclic ether)

### 18.1 Nomenclature of Ethers

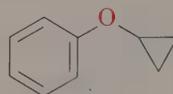
Two different systems for naming ethers are allowed by IUPAC rules. Relatively simple ethers are named by identifying the two organic residues and adding the word *ether*. For example:



*tert*-Butyl methyl ether

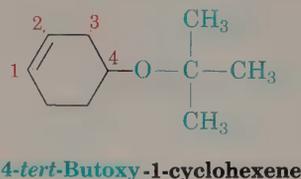
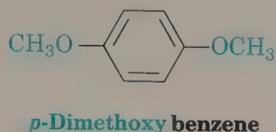


Ethyl vinyl ether



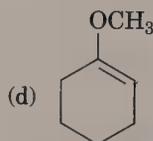
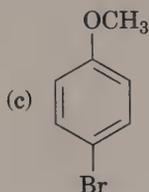
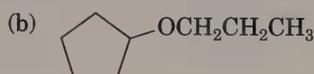
Cyclopropyl phenyl ether

If more than one ether linkage is present in the molecule, or if other functional groups are present, the ether is named as an alkoxy-substituted parent compound. For example:



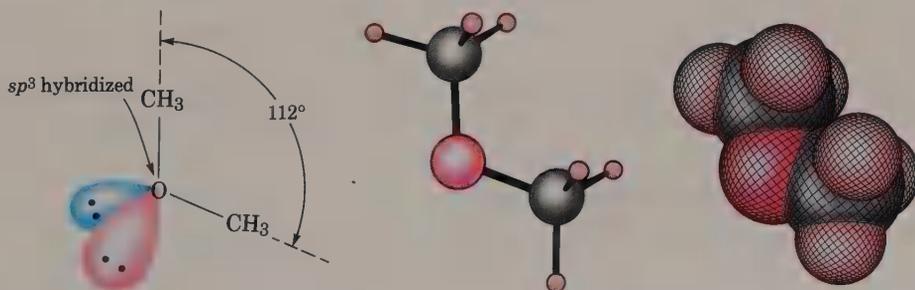
PROBLEM .....

18.1 Name these ethers according to IUPAC rules:



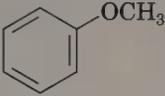
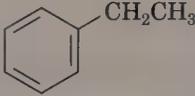
## 18.2 Structure and Properties of Ethers

Ethers can be thought of as organic derivatives of water in which the hydrogen atoms have been replaced by organic fragments,  $\text{H}-\text{O}-\text{H}$  versus  $\text{R}-\text{O}-\text{R}$ . As such, ethers have nearly the same geometry as water. The  $\text{R}-\text{O}-\text{R}$  bonds have an approximately tetrahedral bond angle ( $112^\circ$  in dimethyl ether), and the oxygen atom is  $sp^3$  hybridized.



The presence of the electronegative oxygen atom causes ethers to have a slight dipole moment, and the boiling points of ethers are therefore somewhat higher than the boiling points of comparable alkanes. Table 18.1 compares the boiling points of some common ethers with the corresponding

TABLE 18.1 Comparison of Boiling Points of Ethers and Hydrocarbons

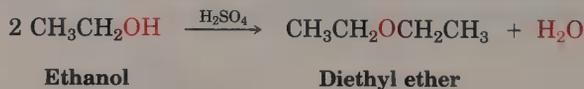
<i>Ether</i>	<i>[Hydrocarbon]</i>	<i>Boiling point (°C)</i>	
CH <sub>3</sub> OCH <sub>3</sub>	CH <sub>3</sub> CH <sub>2</sub> CH <sub>3</sub>	-25	-45
CH <sub>3</sub> CH <sub>2</sub> OCH <sub>2</sub> CH <sub>3</sub>	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>3</sub>	34.6	36
		65	49
		158	136

hydrocarbons in which the ether oxygen atom has been replaced by a CH<sub>2</sub> group.

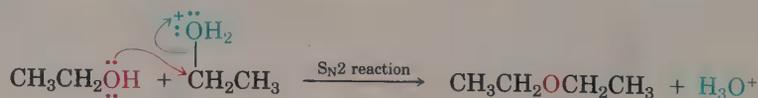
Although ethers are inert toward many reagents, certain ethers react slowly with air to give **peroxides**, compounds that contain oxygen–oxygen bonds. The peroxides from low-molecular-weight ethers such as diisopropyl ether and tetrahydrofuran are highly explosive and extremely dangerous, even in tiny amounts. This sensitivity to air, combined with their volatility and flammability, makes ether solvents potentially dangerous unless handled by skilled persons. Ether solvents are very useful in the laboratory, but they must always be treated with care.

### 18.3 Industrial Preparation of Ethers

Diethyl ether and other simple symmetrical ethers are prepared industrially by the sulfuric acid–catalyzed dehydration of alcohols:



The reaction probably occurs by S<sub>N</sub>2 displacement of water from a protonated ethanol molecule by the oxygen atom of a second ethanol:



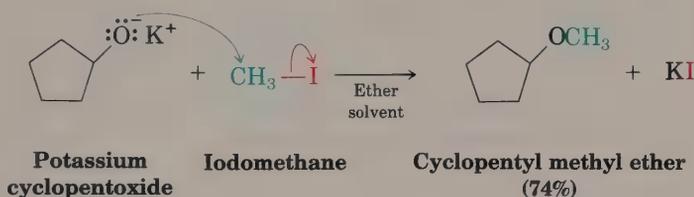
This acid-catalyzed method of ether preparation is limited to the industrial synthesis of symmetrical ethers from primary alcohols, since secondary and tertiary alcohols dehydrate to yield alkenes. The reaction conditions must be carefully controlled, and the method is of little practical value in the laboratory.

## PROBLEM

- 18.2 Why do you suppose only symmetrical ethers can be prepared by the sulfuric acid-catalyzed dehydration procedure? What product(s) would you expect to get if ethanol and 1-propanol were allowed to react together? In what ratio would the products be formed if the two alcohols were of equal reactivity?

## 18.4 The Williamson Ether Synthesis

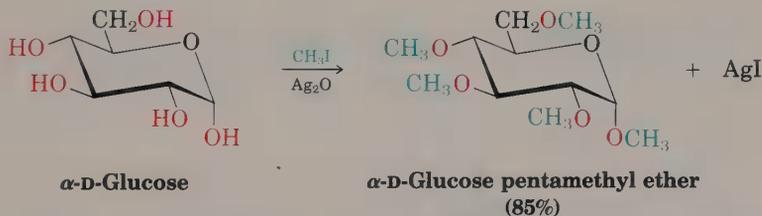
Metal alkoxides react with primary alkyl halides and tosylates by an  $S_N2$  pathway to yield ethers, a process known as the **Williamson ether synthesis**. Discovered in 1850, the Williamson<sup>1</sup> synthesis is still the best method for the preparation of ethers, both symmetrical and unsymmetrical.



The alkoxide ions needed in the Williamson reaction are normally prepared by reaction of an alcohol with a strong base such as sodium hydride, NaH (Section 17.4). An acid–base reaction occurs between the alcohol and sodium hydride to generate the sodium salt of the alcohol.



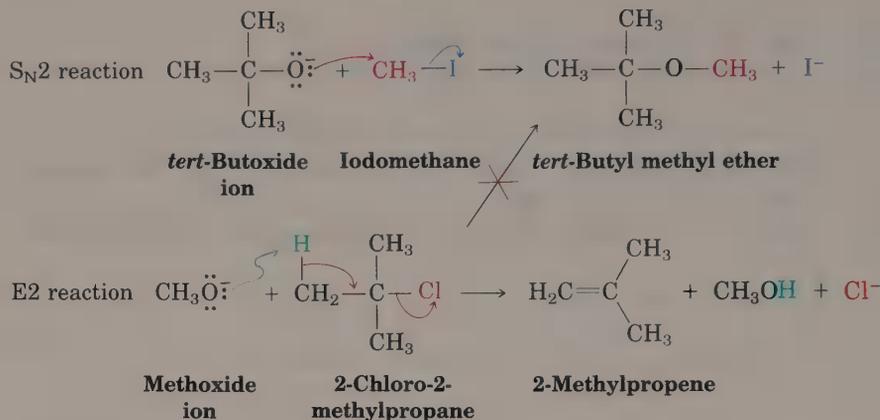
An important variation of the Williamson synthesis involves the use of silver oxide,  $\text{Ag}_2\text{O}$ , as base, rather than NaH. Under these conditions, the free alcohol reacts directly with alkyl halide, and there is no need to preform the metal alkoxide salt. For example, glucose reacts with iodomethane in the presence of  $\text{Ag}_2\text{O}$  to generate a *pentaether* in 85% yield.



Mechanistically, the Williamson synthesis is simply an  $S_N2$  displacement of halide ion by an alkoxide ion nucleophile. The Williamson synthesis is thus subject to all of the normal constraints on  $S_N2$  reactions discussed

<sup>1</sup>Alexander W. Williamson (1824–1904); b. Wandsworth, England; Ph.D. Giessen (1846); professor, University College, London (1849–1904).

in Section 11.5. Primary halides and tosylates work best, since competitive E2 elimination of HX can occur with more hindered substrates. For this reason, unsymmetrical ethers should be synthesized by reaction between the more hindered alkoxide partner and less hindered halide partner, rather than vice versa. For example, *tert*-butyl methyl ether is best prepared by reaction of *tert*-butoxide ion with iodomethane, rather than by reaction of methoxide ion with 2-chloro-2-methylpropane:



PROBLEM .....

- 18.3 Treatment of cyclohexanol with NaH gives an alkoxide ion that reacts with iodomethane to yield an ether. Write out the reaction showing all steps.

PROBLEM .....

- 18.4 How would you prepare these compounds using a Williamson synthesis?
- |                            |                                    |
|----------------------------|------------------------------------|
| (a) Methyl propyl ether    | (b) Anisole (methyl phenyl ether)  |
| (c) Benzyl isopropyl ether | (d) Ethyl 2,2-dimethylpropyl ether |

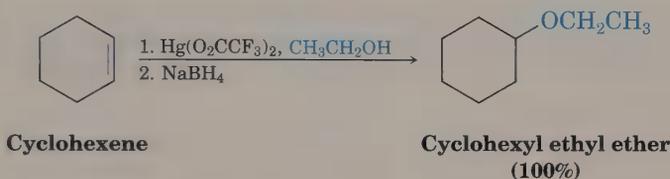
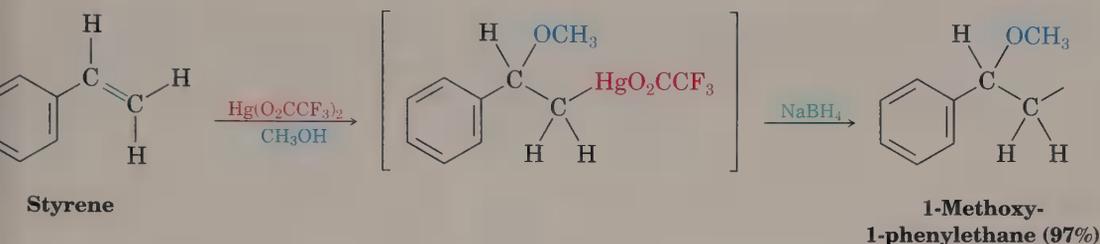
PROBLEM .....

- 18.5 Rank the following halides in order of their expected reactivity in the Williamson synthesis.
- |   |
|---|
| (a) Bromoethane, 2-bromopropane, bromobenzene |
| (b) Chloroethane, bromoethane, 1-iodopropene  |
- .....

## 18.5 Alkoxymercuration–Demercuration of Alkenes

We saw in Section 7.4 that alkenes react with water in the presence of mercuric acetate to yield a hydroxymercuration product. Subsequent treatment of this addition product with sodium borohydride breaks the carbon–

mercury bond and yields the alcohol. A similar **alkoxymercuration** reaction occurs when an alkene is treated with an *alcohol* in the presence of mercuric acetate. [Mercuric trifluoroacetate,  $\text{Hg}(\text{OOCF}_3)_2$ , works even better.] Sodium borohydride–induced demercuration then yields an ether. As indicated by the following examples, the net result is Markovnikov addition of alcohol to the alkene:



The mechanism of the alkoxymercuration–demercuration reaction is analogous to the mechanism described earlier (Section 7.4) for the hydroxymercuration procedure. The reaction is initiated by electrophilic addition of mercuric ion to the alkene, followed by reaction of the intermediate cation with alcohol. Displacement of mercury with sodium borohydride completes the process.

A wide variety of alcohols and alkenes can be used in the alkoxymercuration reaction. Primary, secondary, and even tertiary alcohols react smoothly, but ditertiary ethers can't be prepared because of steric hindrance to reaction.

PROBLEM .....

- 18.6 Show in detail the mechanism of the reaction of 1-methylcyclopentene with ethanol and mercuric trifluoroacetate.

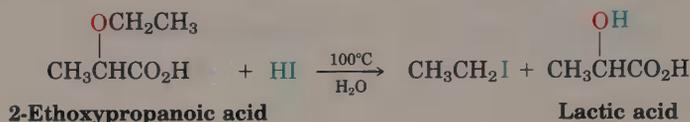
PROBLEM .....

- 18.7 How would you prepare the following ethers? Use whichever method you think is most appropriate, the Williamson synthesis or the alkoxymercuration reaction.
- Butyl cyclohexyl ether
  - Benzyl ethyl ether ( $\text{CH}_3\text{CH}_2\text{OCH}_2\text{C}_6\text{H}_5$ )
  - tert*-Butyl *sec*-butyl ether
  - Tetrahydrofuran
- .....

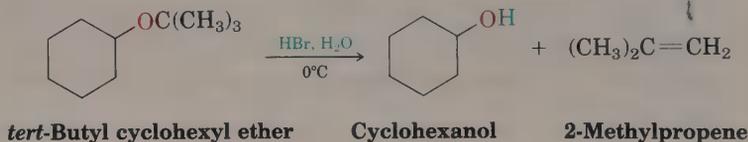
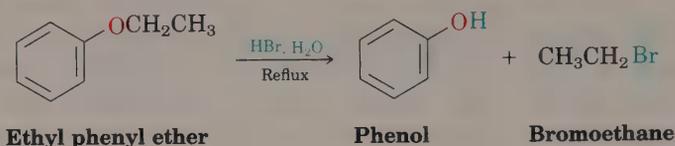
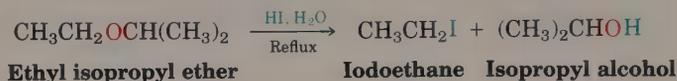
## 18.6 Reactions of Ethers: Acidic Cleavage

Ethers are unusually inert to most reagents used in organic chemistry, a property that accounts for their wide use as reaction solvents. Halogens, mild acids, bases, and nucleophiles have no effect on most ethers. In fact, ethers undergo only one reaction of general use: They are cleaved by strong acids.

The first example of acid-induced ether cleavage was observed in 1861 by Alexander Butleroff,<sup>2</sup> who found that 2-ethoxypropanoic acid reacts with aqueous HI at 100°C to yield iodoethane and lactic acid:

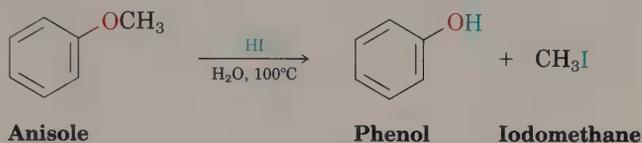
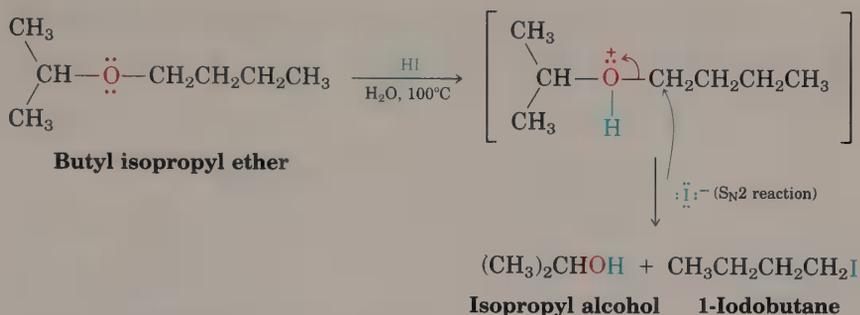


Aqueous HI is still the preferred reagent for cleaving simple ethers, although HBr can also be used. HCl does not cleave ethers.

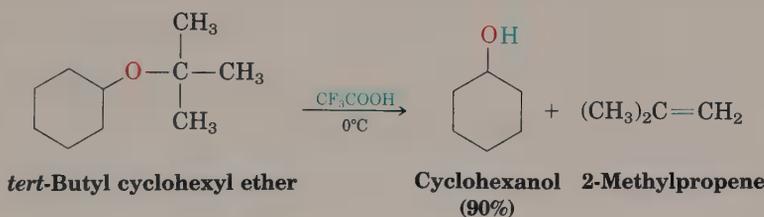


Acidic ether cleavages are typical nucleophilic substitution reactions that take place by either an S<sub>N</sub>1 pathway or an S<sub>N</sub>2 pathway, depending on the structure of the ether. Primary and secondary alkyl ethers react by an S<sub>N</sub>2 pathway in which iodide or bromide ion attacks the protonated ether at the less highly substituted site. This usually results in a selective cleavage into a single alcohol and a single alkyl halide, rather than a mixture of products. For example, butyl isopropyl ether yields exclusively isopropyl alcohol and 1-iodobutane on cleavage by HI, since nucleophilic attack by iodide ion occurs at the less hindered primary site rather than at the more hindered secondary site. Similarly, anisole is cleaved to give phenol and iodomethane.

<sup>2</sup>Alexander M. Butleroff (1828–1886); b. Tschistopol, Russia; Ph.D., University of Moscow (1854); professor, University of Kazan (1854–1867), University of St. Petersburg (1867–1880).



Tertiary, benzylic, and allylic ethers tend to cleave by either an  $\text{S}_{\text{N}}1$  or an  $\text{E}1$  mechanism, since these substrates are capable of producing stabilized intermediate carbocations. These reactions are often fast and take place at moderate temperatures. *tert*-Butyl ethers, for example, can often be cleaved at room temperature or below. Trifluoroacetic acid, rather than  $\text{HBr}$  or  $\text{HI}$ , seems to be best for these reactions, as the following example shows:



PROBLEM .....

- 18.8 Write the mechanism of the trifluoroacetic acid-catalyzed cleavage of a *tert*-butyl ether. Account for the fact that 2-methylpropene is formed.

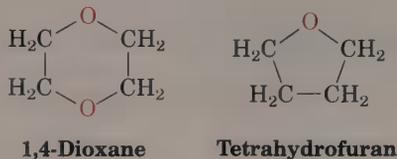
PROBLEM .....

- 18.9 Suggest an explanation for the observation that  $\text{HI}$  and  $\text{HBr}$  are much more effective than  $\text{HCl}$  in cleaving ethers. [Hint: See Section 11.5.]
- .....

## 18.7 Cyclic Ethers: Epoxides

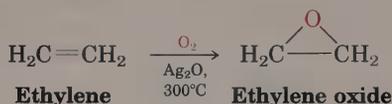
For the most part, cyclic ethers behave like acyclic ethers. The chemistry of the ether functional group is the same, whether it's in an open chain or in a ring. For example, common cyclic ethers such as tetrahydrofuran and

dioxane are often used as solvents because of their inertness, yet they can be cleaved by strong acids.

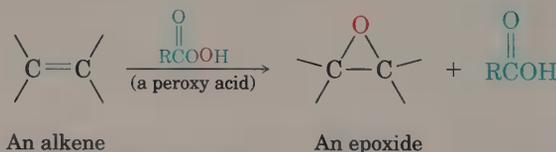


The only group of cyclic ethers that behaves differently from open-chain ethers is the group of three-membered ring compounds called **epoxides**, or **oxiranes**. The strain of the three-membered ether ring makes epoxides highly reactive and gives them unique chemical reactivity.

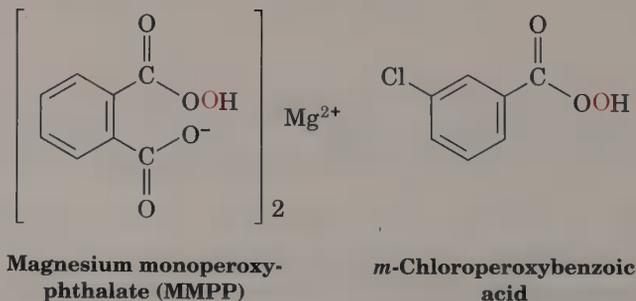
Ethylene oxide, the simplest epoxide, is an intermediate in the manufacture of both ethylene glycol for automobile antifreeze and polyester polymers. More than 2.7 million tons of ethylene oxide are produced industrially each year by air oxidation of ethylene over a silver oxide catalyst at 300°C. This process is not of general utility for other epoxides, however, and is of little value in the laboratory.

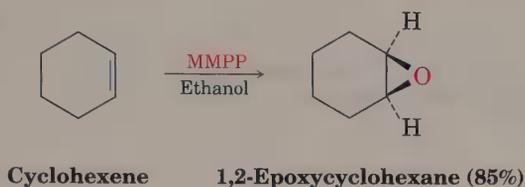


In the laboratory, epoxides are normally prepared by treatment of an alkene with a **peroxyacid**,  $\text{RCO}_3\text{H}$ :

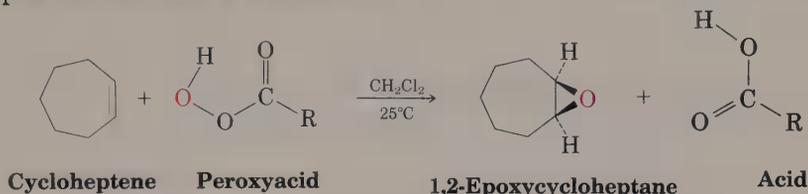


Many different peroxyacids can be used to accomplish epoxidation. On a laboratory scale, *m*-chloroperoxybenzoic acid was the preferred reagent for many years, but recent concerns about safety have caused it to be replaced by the more stable magnesium monoperoxyphthalate (MMPP).

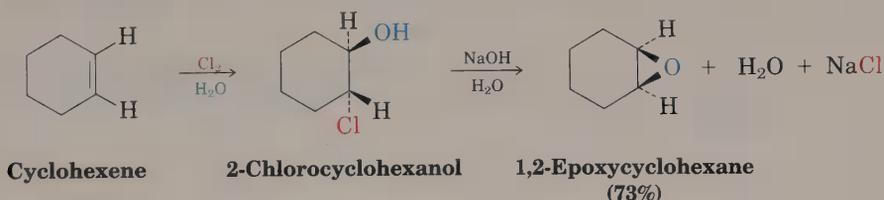




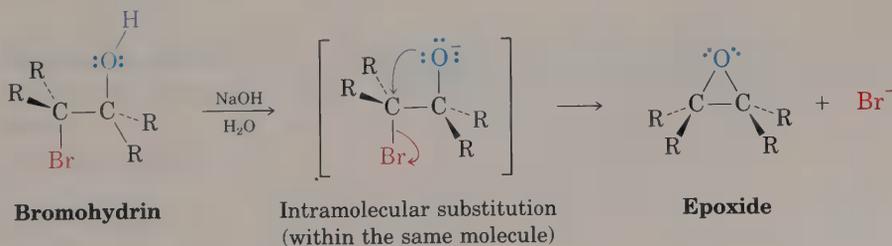
Peroxyacids transfer oxygen to an alkene through a one-step mechanism without intermediates. Thus, the epoxidation reaction is different from other alkene addition reactions we've studied. No carbocation intermediate is involved, and the details of the process aren't fully understood. There is good evidence to show, however, that the oxygen farthest from the carbonyl group is the one that is transferred.



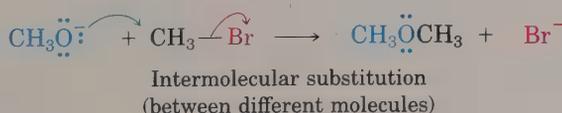
Another method for the synthesis of epoxides is through the use of halohydrins, prepared by electrophilic addition of HO-X to alkenes (Section 7.3). When halohydrins are treated with base, HX is eliminated, and an epoxide is produced.



Note that the synthesis of epoxides by base treatment of halohydrins is actually an *intramolecular* Williamson ether synthesis. The nucleophilic oxygen atom and the electrophilic carbon atom are in the same molecule, rather than in different molecules:

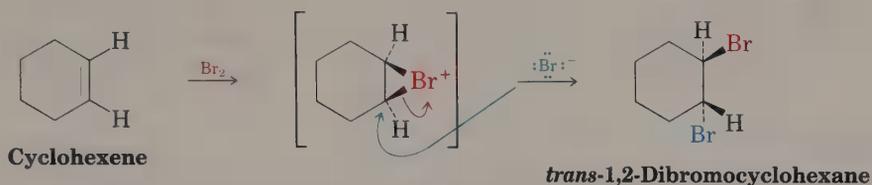


Recall the following:

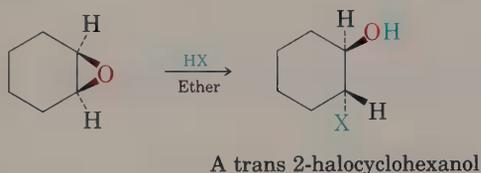




Recall the following:



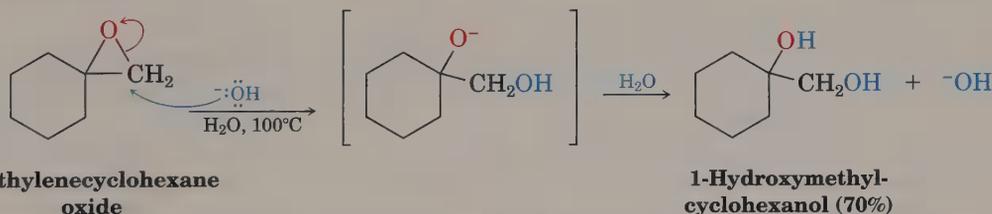
Protonated epoxides can also be opened by nucleophiles other than water. For example, if anhydrous HX is used, epoxides can be converted into *trans* halohydrins:



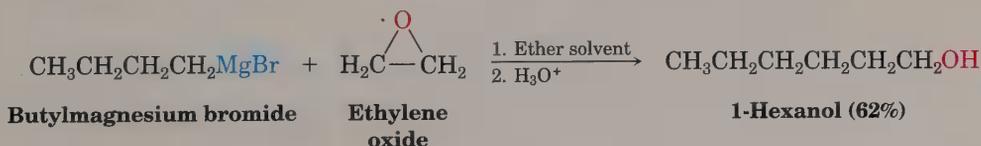
where X = F, Br, Cl, or I

Unlike other ethers, epoxides can also be cleaved by base and by other nucleophiles. Although an ether oxygen is normally a very poor leaving group in an  $\text{S}_{\text{N}}2$  reaction (Section 11.5), the reactivity of the three-membered ring is sufficient to cause epoxides to react with hydroxide ion at elevated temperatures.

$\text{S}_{\text{N}}2$  reaction



A similar nucleophilic ring opening is found when epoxides are treated with Grignard reagents. Ethylene oxide is frequently used, allowing the conversion of a Grignard reagent into a primary alcohol having two more carbons than the starting alkyl halide. 1-Bromobutane is converted into 1-hexanol, for example, by reaction of its Grignard reagent with ethylene oxide:



PROBLEM 18.12

18.12 Show all of the steps involved in the acidic hydrolysis of *cis*-5,6-epoxydecane. What is the stereochemistry of the product, assuming normal back-side  $\text{S}_{\text{N}}2$  attack?

## PROBLEM .....

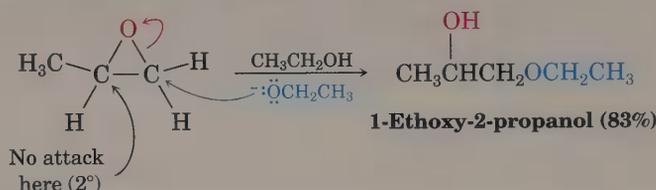
- 18.13 What is the stereochemistry of the product from acid-catalyzed hydrolysis of *trans*-5,6-epoxydecane? Is the product the same as or different from the one formed in Problem 18.12?

## PROBLEM .....

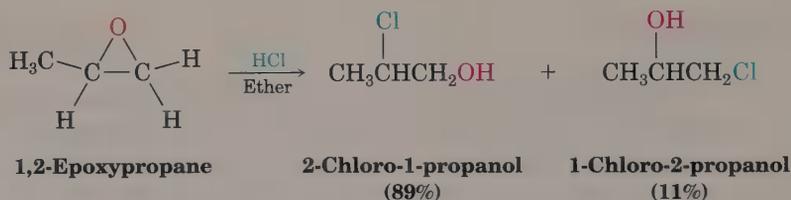
- 18.14 Acid-induced hydrolysis of a 1,2-epoxycyclohexane produces a *trans*-diaxial 1,2-diol. What product would you expect to obtain from acidic hydrolysis of *cis*-3-*tert*-butyl-1,2-epoxycyclohexane? (Recall that the bulky *tert*-butyl group locks the cyclohexane ring into a specific conformation.)

## REGIOCHEMISTRY OF EPOXIDE RING OPENING

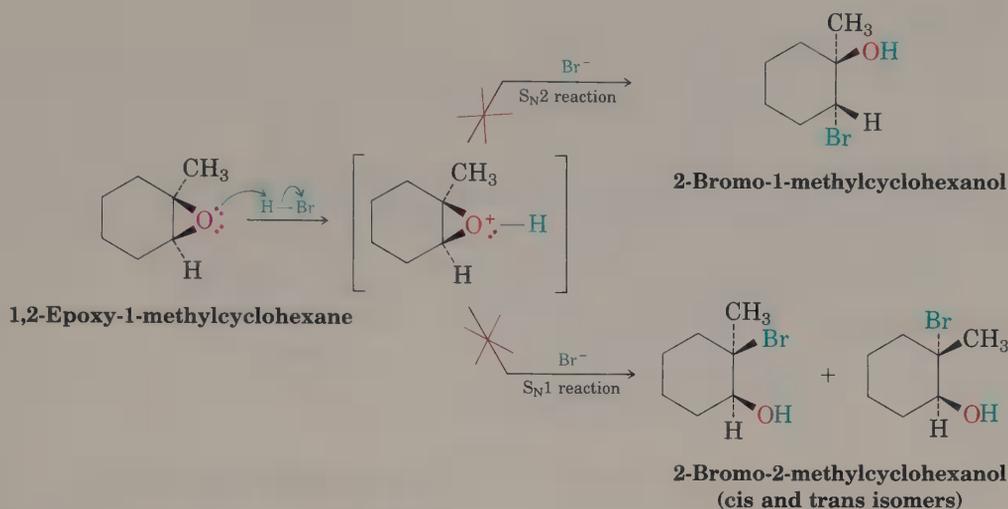
The direction in which an unsymmetrical epoxide ring is opened depends on the conditions used. If a basic nucleophile is used in a typical  $S_N2$ -type reaction, attack takes place at the less hindered epoxide carbon. For example, 1,2-epoxypropane is attacked by ethoxide ion exclusively at the less highly substituted, primary carbon to give 1-ethoxy-2-propanol.



If acidic conditions are used, however, a different reaction course is followed, and attack of the nucleophile occurs primarily at the *more* highly substituted carbon atom.



Acid-catalyzed epoxide opening is particularly interesting from a mechanistic viewpoint because it appears to be *midway* between a pure  $S_N1$  reaction and a pure  $S_N2$  reaction. Take the reaction of 1,2-epoxy-1-methylcyclohexane with HBr, for example. If this were a pure  $S_N2$  reaction, bromide ion would attack the less highly substituted carbon atom, displacing the epoxide oxygen atom from the back side and yielding 2-bromo-1-methylcyclohexanol with the  $-\text{Br}$  and  $-\text{OH}$  groups *trans* to each other. On the other hand, if this were a pure  $S_N1$  reaction, protonation of the epoxide oxygen atom followed by ring opening would yield a carbocation that could react with bromide ion from either side to give a mixture of two isomeric 2-bromo-2-methylcyclohexanols. One isomer would have the  $-\text{Br}$  and  $-\text{OH}$  groups *cis*; the other would have  $-\text{Br}$  and  $-\text{OH}$  *trans* (Figure 18.1).



**Figure 18.1** Reaction products from hypothetical  $S_N1$  and  $S_N2$  ring openings of 1,2-epoxy-1-methylcyclohexane. In practice, neither reaction course is followed.

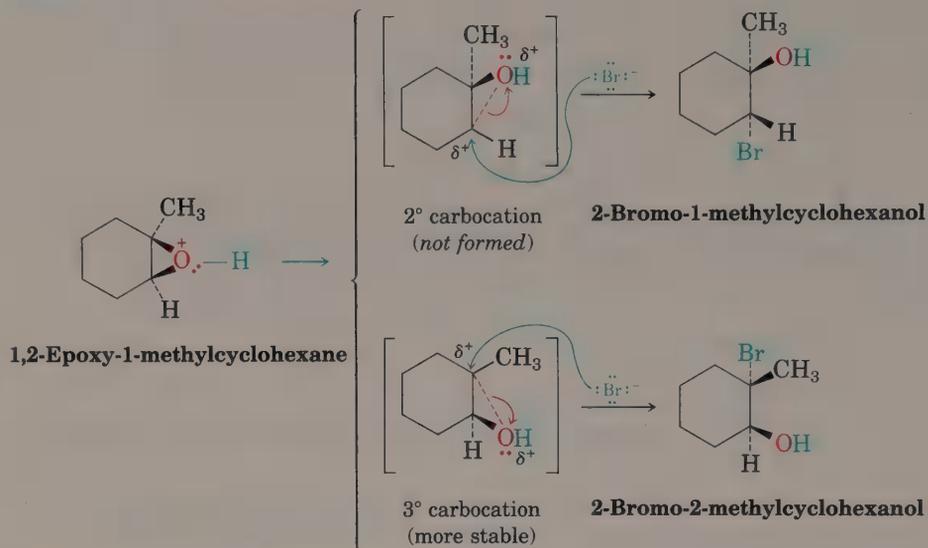
In fact, *neither* of the reaction courses in Figure 18.1 is followed. Instead, reaction of 1,2-epoxy-1-methylcyclohexane with HBr yields a single isomer of 2-bromo-2-methylcyclohexanol in which the bromo and hydroxyl groups are trans to each other. The observed regiochemistry can be accounted for by assuming that the reaction has characteristics of *both*  $S_N1$  and  $S_N2$  reactions. The fact that the single product formed has the entering bromine and the leaving oxygen on opposite sides of the ring is clearly an  $S_N2$ -like result with back-side displacement of the leaving group. The fact that bromide ion attacks the tertiary side of the epoxide rather than the secondary side is clearly an  $S_N1$ -like result.

Both facts can be accommodated by postulating that the transition state for acid-induced epoxide ring opening has an  $S_N2$ -like geometry but a high degree of  $S_N1$ -like carbocationic character. Although the protonated epoxide is not a full carbocation, it is strongly polarized so that the positive charge is shared by the more highly substituted carbon atom. Thus, attack of a nucleophile occurs at the more highly substituted site (Figure 18.2).

**PROBLEM** .....

**18.15** Predict the major product of each of these reactions:

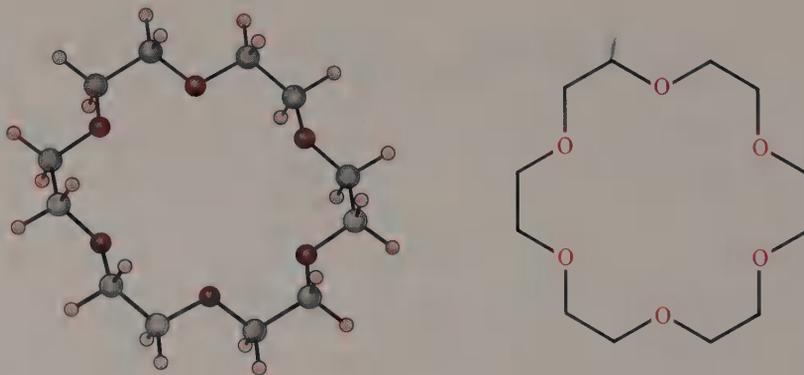




**Figure 18.2** Acid-induced ring opening of 1,2-epoxy-1-methylcyclohexane. There is a high degree of  $S_N1$ -like carbocationic character in the transition state, leading to exclusive formation of the isomer of 2-bromo-2-methylcyclohexanol that has  $-Br$  and  $-OH$  groups trans.

## 18.9 Crown Ethers

**Crown ethers**, discovered in the early 1960s by Charles Pedersen<sup>3</sup> at the Du Pont Company, are a relatively recent addition to the ether family. Crown ethers are named according to the general format  $x$ -crown- $y$ , where  $x$  is the total number of atoms in the ring and  $y$  is the number of oxygen atoms. Thus, 18-crown-6 ether is an 18-membered ring containing 6 ether oxygen atoms.



18-Crown-6 ether

<sup>3</sup>Charles John Pedersen (1904–1989); b. Pusan, Korea (U.S. citizen); M.Sc. Massachusetts Institute of Technology (1927); E. I. du Pont de Nemours & Co. (1927–1969); Nobel Prize (1987).

The importance of crown ethers derives from their extraordinary ability to solvate metal cations by sequestering the metal in the center of the polyether cavity. Different crown ethers solvate different metal cations, depending on the match between ion size and cavity size. For example, 18-crown-6 is able to complex strongly with potassium ion.

Complexes between crown ethers and inorganic salts are soluble in nonpolar organic solvents, thus allowing many reactions to be carried out under aprotic conditions that would otherwise have to be carried out in aqueous solution. For example, the inorganic salt  $\text{KMnO}_4$  actually dissolves in benzene in the presence of 18-crown-6. The resulting solution of “purple benzene” is a valuable reagent for oxidizing alkenes.



Many other inorganic salts, including  $\text{KF}$ ,  $\text{KCN}$ , and  $\text{NaN}_3$ , can be made soluble in organic solvents by the use of crown ethers. The effect of using a crown ether to dissolve a salt in a hydrocarbon or ether solvent is similar to the effect of dissolving the salt in a polar aprotic solvent such as DMSO, DMF, or HMPA (Section 11.5). In both cases, the metal cation is strongly solvated, leaving the anion bare. Thus, the  $\text{S}_{\text{N}}2$  reactivity of an anion is tremendously enhanced in the presence of a crown ether.

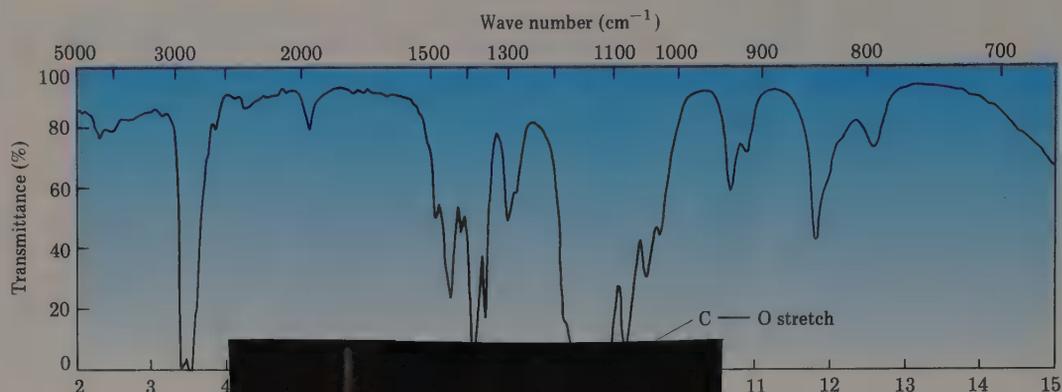
PROBLEM.....

- 18.16 15-Crown-5 and 12-crown-4 ethers complex  $\text{Na}^+$  and  $\text{Li}^+$ , respectively. Make models of these crown ethers, and compare the sizes of the cavities.
- .....

## 18.10 Spectroscopic Analysis of Ethers

### INFRARED SPECTROSCOPY

Ethers are difficult to distinguish by infrared spectroscopy. Although they show an absorption due to carbon–oxygen single-bond stretching in the range  $1050\text{--}1150\text{ cm}^{-1}$ , many other kinds of absorptions occur in the same range. Figure 18.3 shows the infrared spectrum of diethyl ether and identifies the C–O stretch.

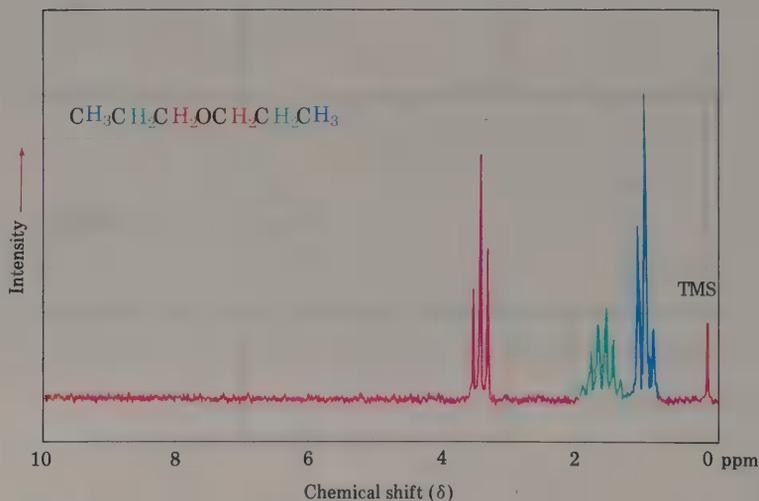


**Figure 18.3** The infrared spectrum of dipropyl ether,  $\text{CH}_3\text{CH}_2\text{OCH}_2\text{CH}_3$ .

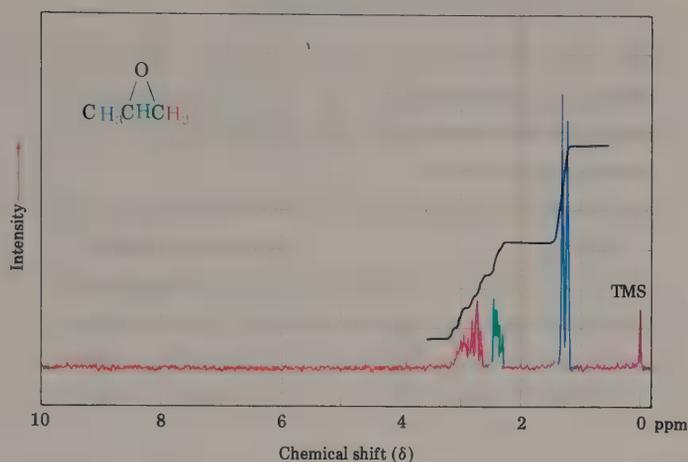
## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Protons on carbon next to oxygen are shifted downfield from the normal alkane resonance and show  $^1\text{H}$  NMR absorptions in the region 3.4–4.5  $\delta$ . This downfield shift is clearly indicated in the spectrum of dipropyl ether shown in Figure 18.4.

Epoxides absorb at a slightly higher field than other ethers and show characteristic resonances at 2.5–3.5  $\delta$  in the  $^1\text{H}$  NMR, as indicated for 1,2-epoxypropane in Figure 18.5.

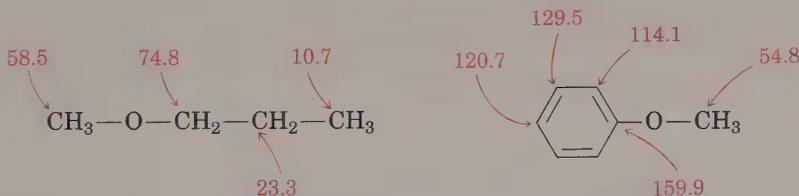


**Figure 18.4** The proton NMR spectrum of dipropyl ether. Protons on carbon next to oxygen are shifted downfield to 3.4  $\delta$ .



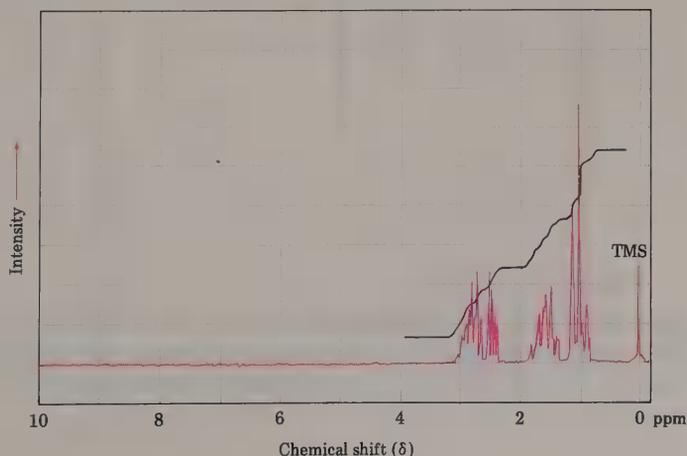
**Figure 18.5** The proton NMR spectrum of 1,2-epoxypropane.

Ether carbon atoms also exhibit a downfield shift in the  $^{13}\text{C}$  NMR spectrum, where they usually absorb in the range 50–80  $\delta$ . For example, the carbon atoms next to oxygen in methyl propyl ether absorb at 58.5 and 74.8  $\delta$ . Similarly, the methyl carbon in anisole absorbs at 54.8  $\delta$ .



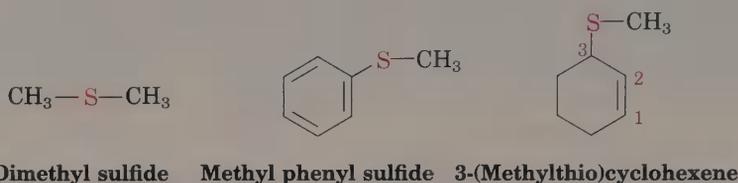
**PROBLEM** .....

**18.17** The  $^1\text{H}$  NMR spectrum shown is that of a substance with the formula  $\text{C}_4\text{H}_8\text{O}$ . Propose a structure.

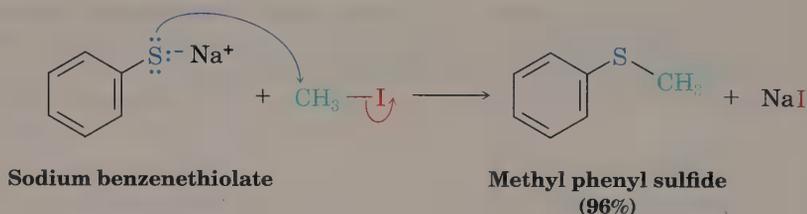


## 18.11 Sulfides

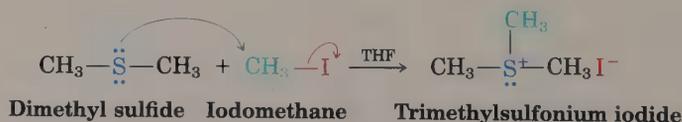
**Sulfides,  $\text{RSR}'$** , are sulfur analogs of ethers. They are named by following the same rules used for ethers, with *sulfide* used in place of *ether* for simple compounds and *alkylthio* used in place of *alkoxy* for more complex substances.



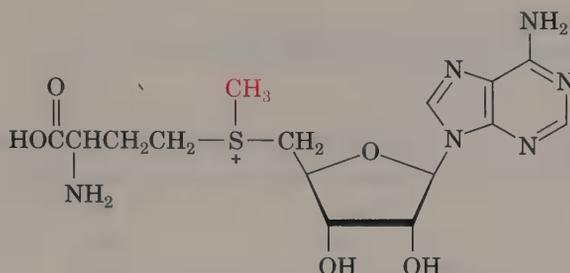
Sulfides are prepared by treatment of a primary or secondary alkyl halide with a **thiolate** anion,  $\text{RS}^-$ . The reaction occurs by an  $\text{S}_{\text{N}}2$  mechanism, analogous to the Williamson synthesis of ethers (Section 18.4). Thiolate anions are among the best nucleophiles known (Section 11.5), and product yields are usually high in these substitution reactions.



Since the valence electrons on sulfur are farther from the nucleus and are less tightly held than those on oxygen (3p electrons versus 2p electrons), there are some important differences between the chemistries of ethers and sulfides. For example, sulfur is more polarizable than oxygen, and sulfur compounds are thus more nucleophilic than their oxygen analogs. Unlike dialkyl ethers, dialkyl sulfides are good nucleophiles that react rapidly with primary alkyl halides by an  $\text{S}_{\text{N}}2$  mechanism. The products of such reactions are **trialkylsulfonium salts ( $\text{R}_3\text{S}^+$ )**.

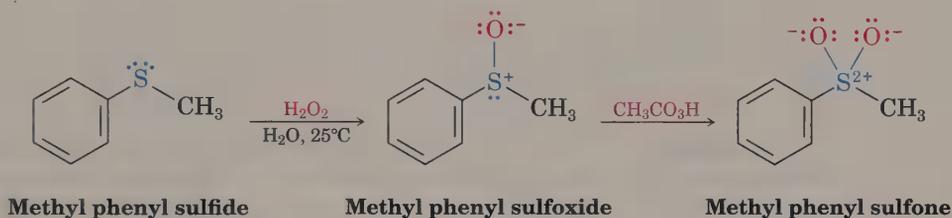


Once formed, trialkylsulfonium salts are themselves good alkylating agents, since a nucleophile can attack one of the groups bound to the positively charged sulfur, displacing a sulfide as leaving group. Nature makes extensive use of the trialkylsulfonium salt *S*-adenosylmethionine, as a biological methylating agent (Section 11.17).

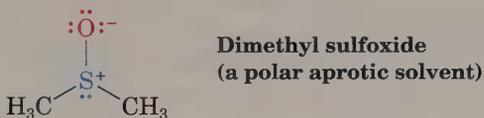


S-Adenosylmethionine (a sulfonium salt)

A second difference between sulfides and ethers is that sulfides are easily oxidized. Treatment of a sulfide with hydrogen peroxide,  $\text{H}_2\text{O}_2$ , at room temperature yields the corresponding **sulfoxide** ( $\text{R}_2\text{SO}$ ), and further oxidation of the sulfoxide with a peroxyacid yields a **sulfone** ( $\text{R}_2\text{SO}_2$ ).

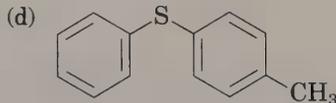
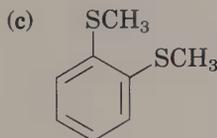
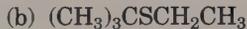
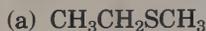


Dimethyl sulfoxide (DMSO) is a particularly well known sulfoxide that finds wide use as a polar aprotic solvent. It must be handled with care, however, because it has a remarkable ability to penetrate the skin, carrying along whatever is dissolved in it.



**PROBLEM** .....

18.18 Name the following compounds by IUPAC rules:



**PROBLEM** .....

18.19 How can you account for the fact that dimethyl sulfoxide has a boiling point of  $189^\circ\text{C}$  and is miscible with water, whereas dimethyl sulfide has a boiling point of  $37^\circ\text{C}$  and is immiscible with water?

.....

## 18.12 Summary and Key Words

**Ethers** are compounds that have two organic groups bonded to the same oxygen atom, ROR'. The organic groups may be alkyl, vinylic, or aryl, and the oxygen atom may be in a ring or in an open chain.

Ethers are normally prepared in the laboratory either by a Williamson synthesis or by an alkoxymercuration–demercuration sequence. The **Williamson ether synthesis** involves S<sub>N</sub>2 attack of an alkoxide ion on a primary alkyl halide. The reaction is thus limited to the preparation of ethers in which one organic group is primary. The **alkoxymercuration–demercuration** sequence involves the formation of an intermediate alkoxy organomercury compound, followed by sodium borohydride reduction of the carbon–mercury bond. The net result is Markovnikov addition of an alcohol to an alkene.

Ethers are inert to most common reagents but are attacked by strong acids to give cleavage products. Both HI and HBr are often used. The cleavage reaction takes place by an S<sub>N</sub>2 mechanism if primary and secondary alkyl groups are bonded to the ether oxygen, but by an S<sub>N</sub>1 mechanism if one of the alkyl groups bonded to oxygen is tertiary.

**Epoxides**, cyclic ethers with a three-membered oxygen-containing ring, differ from other ethers in their ease of cleavage. The high reactivity of the three-membered ether ring because of ring strain allows epoxide rings to be opened by nucleophilic attack of strong bases as well as by acids. Base-catalyzed epoxide ring opening occurs by S<sub>N</sub>2 attack of a nucleophile at the less hindered epoxide carbon, whereas acid-induced epoxide ring opening occurs by S<sub>N</sub>1-like attack at the more highly substituted epoxide carbon.

**Sulfides**, RSR', are sulfur analogs of ethers. They are prepared by a Williamson-type S<sub>N</sub>2 reaction between a thiolate anion and a primary or secondary alkyl halide. Sulfides are much more nucleophilic than ethers and can be oxidized to **sulfoxides** (R<sub>2</sub>SO) and to **sulfones** (R<sub>2</sub>SO<sub>2</sub>). Sulfides can also be alkylated by reaction with a primary alkyl halide to yield **sulfonium salts**, R<sub>3</sub>S<sup>+</sup>.

Although not readily identifiable by infrared spectroscopy, ethers show characteristic downfield <sup>1</sup>H NMR absorptions that are easily detected. Carbon atoms bonded to oxygen show a similar downfield shift in the <sup>13</sup>C NMR.

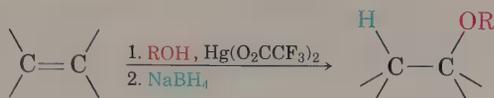
## 18.13 Summary of Reactions

1. Preparation of ethers
  - a. Williamson synthesis (Section 18.4)



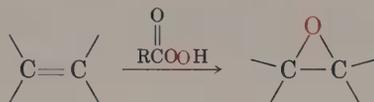
Alkyl halide should be primary.

- b. Alkoxymercuration–demercuration (Section 18.5)



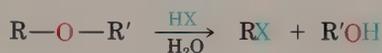
Markovnikov orientation is observed.

- c. Epoxidation of alkenes with peroxyacids (Section 18.7)



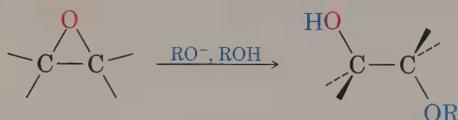
2. Reaction of ethers

- a. Cleavage by HX (Section 18.6)

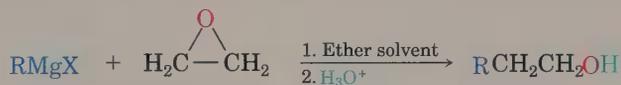


where HX = HBr or HI

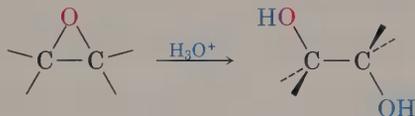
- b. Base-catalyzed epoxide ring opening (Section 18.8)



Reaction occurs at least hindered site.



- c. Acid-catalyzed hydrolysis of epoxides (Section 18.8)



Reaction occurs at more hindered site; trans 1,2-diols are produced from cyclic epoxides.

- d. Acid-induced epoxide ring opening (Section 18.8)



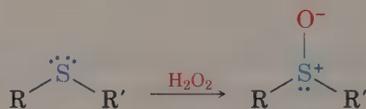
Reaction occurs at more hindered site.

3. Preparation of sulfides (Section 18.11)



## 4. Oxidation of sulfides (Section 18.11)

## a. Preparation of sulfoxides



## b. Preparation of sulfones

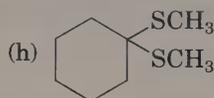
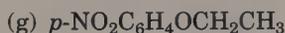
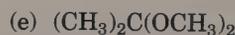
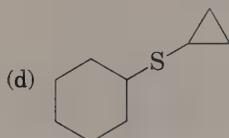
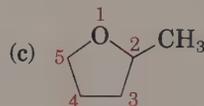
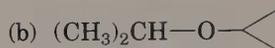
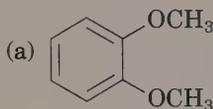


## ADDITIONAL PROBLEMS

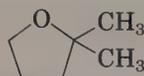
18.20 Draw structures corresponding to these IUPAC names:

- (a) Ethyl 1-ethylpropyl ether                      (b) Di(*p*-chlorophenyl) ether  
 (c) 3,4-Dimethoxybenzoic acid                      (d) Cyclopentylloxycyclohexane  
 (e) 4-Allyl-2-methoxyphenol (eugenol; from oil of cloves)

18.21 Provide correct IUPAC names for these structures:

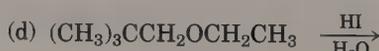
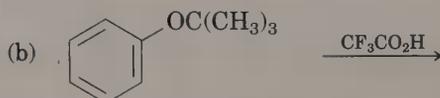
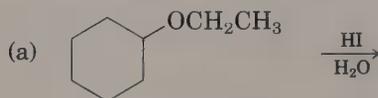


18.22 When 2-methylpentane-2,5-diol is treated with sulfuric acid, dehydration occurs and 2,2-dimethyltetrahydrofuran is formed. Suggest a mechanism for this reaction. Which of the two oxygen atoms is most likely to be eliminated, and why?

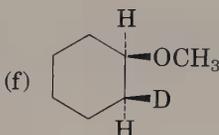
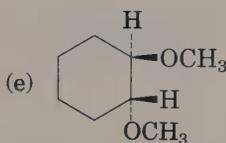
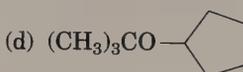
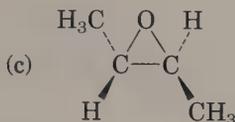
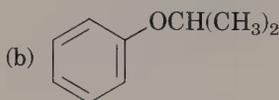
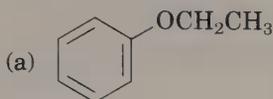


2,2-Dimethyltetrahydrofuran

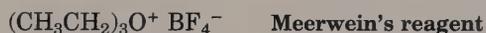
18.23 Predict the products of these ether cleavage reactions:



- 18.24** The *Zeisel method* is an analytical procedure for determining the number of methoxyl groups in a compound. A weighed amount of the compound is heated with concentrated HI, ether cleavage occurs, and the iodomethane product is distilled off and passed into an alcohol solution of  $\text{AgNO}_3$ . The silver iodide that precipitates is then collected and weighed, and the percentage of methoxyl groups in the sample is thereby determined. For example, vanillin, the material responsible for the characteristic odor of vanilla, gives a positive Zeisel reaction; 1.06 g vanillin yields 1.60 g AgI. If vanillin has a molecular weight of 152, how many methoxyls does it contain?
- 18.25** How would you prepare these ethers?

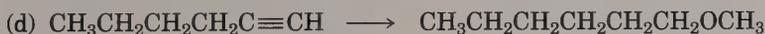
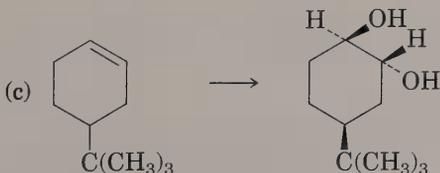
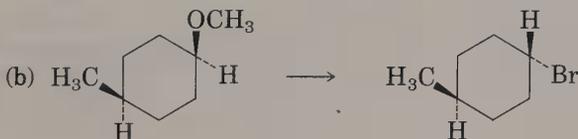
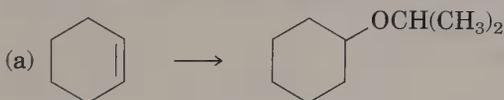


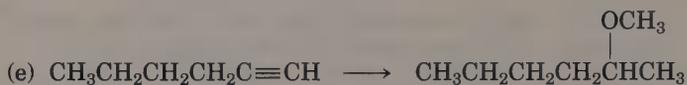
- 18.26** *Meerwein's reagent*, triethyloxonium tetrafluoroborate, is a powerful ethylating agent that converts alcohols into ethyl ethers at neutral pH. Formulate the reaction of Meerwein's reagent with cyclohexanol and account for the fact that trialkyloxonium salts are much more reactive alkylating agents than alkyl iodides are.



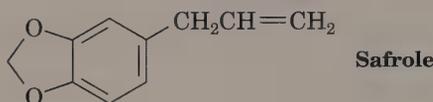
- 18.27** How would you prepare these compounds from 1-phenylethanol?
- (a) Methyl 1-phenylethyl ether                      (b) Phenylepoxyethane  
(c) *tert*-Butyl 1-phenylethyl ether

- 18.28** How would you carry out these transformations? More than one step may be required.

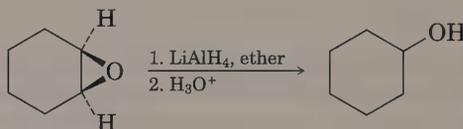




- 18.29 What product(s) would you expect from cleavage of tetrahydrofuran with HI?
- 18.30 How could you prepare benzyl phenyl ether from benzene? More than one step is required.
- 18.31 Methyl aryl ethers can be cleaved to iodomethane and a phenoxide anion by treatment with LiI in hot dimethylformamide solvent. Propose a mechanism for this reaction. What type of mechanism is involved?
- 18.32 *tert*-Butyl ethers can be prepared by the reaction of an alcohol with 2-methylpropene in the presence of an acid catalyst. Propose a mechanism for this reaction.
- 18.33 Safrole, isolated from oil of sassafras, is used as a perfumery agent. Propose a synthesis of safrole from catechol (1,2-benzenediol).



- 18.34 Epoxides are reduced by treatment with lithium aluminum hydride to produce alcohols. Propose a mechanism for this reaction.

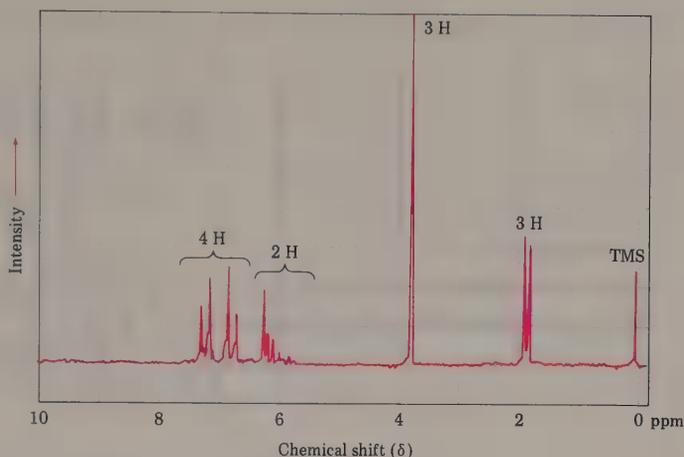


- 18.35 Show the structure and stereochemistry of the alcohol that would result if 1,2-epoxycyclohexane (Problem 18.34) were reduced with lithium aluminum deuteride,  $\text{LiAlD}_4$ .
- 18.36 Grignard reagents react with oxetane to produce primary alcohols, but the reaction is much slower than the corresponding reaction with ethylene oxide. Suggest a reason for the difference in reactivity between oxetane and ethylene oxide.



Oxetane

- 18.37 Treatment of *trans*-2-chlorocyclohexanol with NaOH yields 1,2-epoxycyclohexane, but reaction of the *cis* isomer under the same conditions yields cyclohexanone. Propose mechanisms for both reactions, and explain why the different results are obtained.
- 18.38 Ethyl vinyl ether reacts with ethanol in the presence of an acid catalyst to yield 1,1-diethoxyethane, rather than 1,2-diethoxyethane. How can you account for the observed regioselectivity of addition?
- 18.39 Anethole,  $\text{C}_{10}\text{H}_{12}\text{O}$ , a major constituent of oil of anise, has the  $^1\text{H}$  NMR spectrum shown here. On oxidation with hot sodium dichromate, anethole yields a compound identified as *p*-methoxybenzoic acid. What is the structure of anethole? Assign all peaks in the NMR spectrum, and account for the observed splitting patterns.

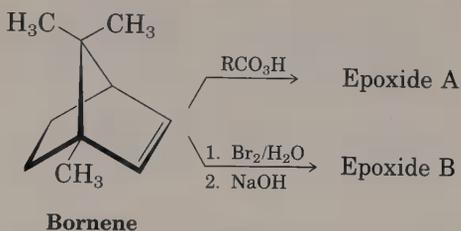


- 18.40** How would you synthesize anethole (Problem 18.39) from benzene?
- 18.41** The red fox (*Vulpes vulpes*) uses a chemical communication system based on scent marks in urine. Recent work has shown one component of fox urine to be a sulfide. Mass spectral analysis of the pure scent-mark component shows  $M^{+} = 116$ . Infrared spectroscopy shows an intense band at  $890\text{ cm}^{-1}$ , and  $^1\text{H}$  NMR spectroscopy reveals the following peaks:

3-proton singlet at  $1.74\ \delta$ ; 3-proton singlet at  $2.11\ \delta$ ; 2-proton triplet at  $2.27\ \delta$ ,  $J = 4.2\ \text{Hz}$ ; 2-proton triplet at  $2.57\ \delta$ ,  $J = 4.2\ \text{Hz}$ ; 2-proton broad peak at  $4.73\ \delta$ .

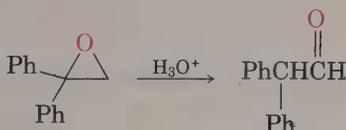
Propose a structure consistent with these data. [Note:  $(\text{CH}_3)_2\text{S}$  absorbs at  $2.1\ \delta$  in the  $^1\text{H}$  NMR.]

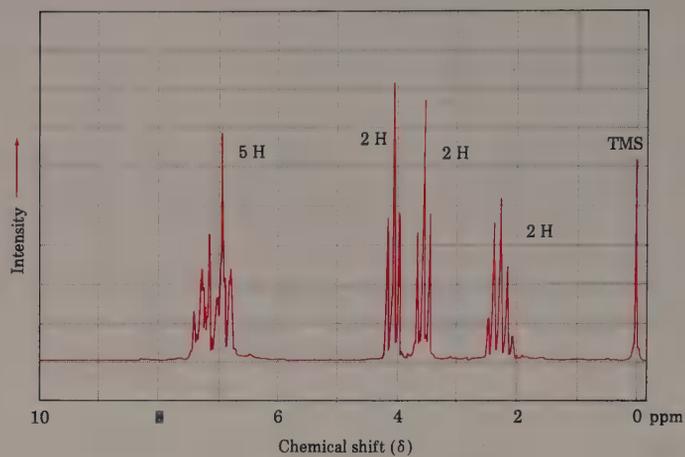
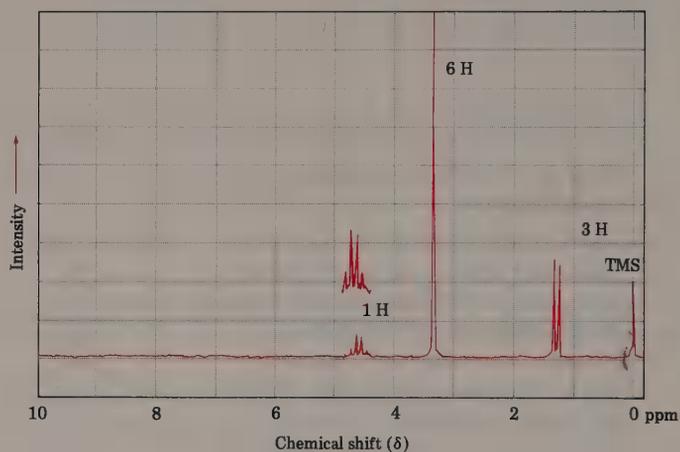
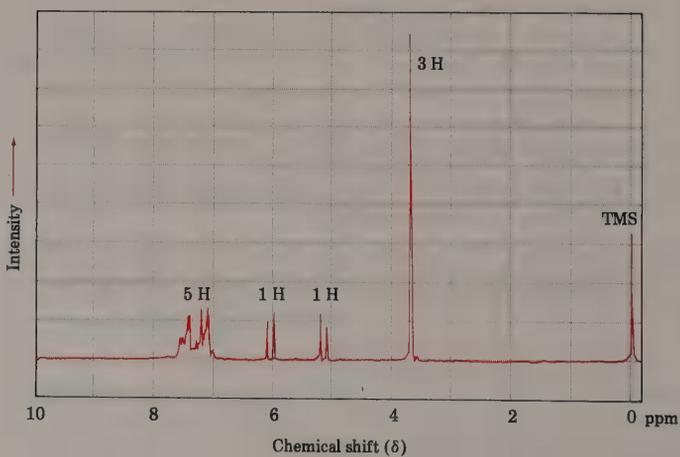
- 18.42** How can you explain the fact that treatment of bornene with a peroxy-acid yields a different epoxide from the one obtained by reaction of bornene with aqueous bromine, followed by base treatment? Propose structures for epoxides A and B.



[Build molecular models of bornene.]

- 18.43** Disparlure,  $\text{C}_{19}\text{H}_{38}\text{O}$ , is a sex attractant released by the female gypsy moth, *Lymantria dispar*. The  $^1\text{H}$  NMR spectrum of disparlure shows a large absorption in the alkane region,  $1\text{--}2\ \delta$ , and a triplet at  $2.8\ \delta$ . Treatment of disparlure, first with aqueous acid and then with  $\text{KMnO}_4$ , yields two carboxylic acids identified as undecanoic acid and 6-methylheptanoic acid. [ $\text{KMnO}_4$  oxidatively cleaves 1,2-diols to yield carboxylic acids (Section 7.10).] Neglecting stereochemistry, propose a structure for disparlure consistent with these data. The actual compound is a chiral molecule with  $7R,8S$  stereochemistry. Draw disparlure, showing the correct stereochemistry.
- 18.44** How would you synthesize racemic disparlure (Problem 18.43) from compounds having 10 or fewer carbons?
- 18.45** Treatment of 1,1-diphenyl-1,2-epoxyethane with aqueous acid yields diphenylacetaldehyde as the major product. Propose a mechanism to account for this reaction.

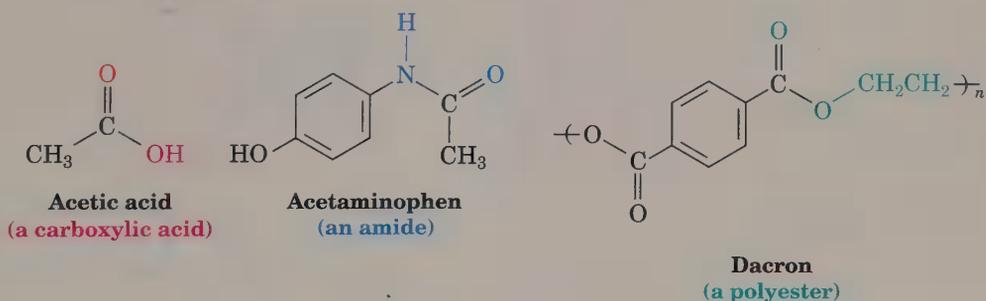


18.46 Propose structures for ethers that have the following  $^1\text{H}$  NMR spectra:(a)  $\text{C}_9\text{H}_{11}\text{BrO}$ (b)  $\text{C}_4\text{H}_{10}\text{O}_2$ (c)  $\text{C}_9\text{H}_{10}\text{O}$ 

# Chemistry of Carbonyl Compounds: An Overview

In this and the next five chapters, we will discuss the most important functional group in organic chemistry: the **carbonyl group**,  $C=O$  (pronounced car-bo-neel).

Carbonyl compounds are everywhere in nature. The majority of biologically important molecules contain carbonyl groups, as do most pharmaceutical agents and many of the synthetic chemicals that touch our everyday lives. Acetic acid, the chief component of vinegar; acetaminophen, the active ingredient in many over-the-counter pain relievers; and Dacron, the polyester material used in clothing, all contain different kinds of carbonyl groups.



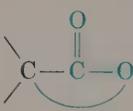
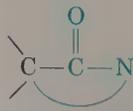
## I. Kinds of Carbonyl Compounds

There are many different kinds of carbonyl compounds, depending on what groups are bonded to the  $C=O$  unit. The chemistry of carbonyl groups is similar, however, regardless of their exact structure.

Table 1 shows some of the many different kinds of carbonyl compounds.

All contain an **acyl fragment**,  $\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-$ , bonded to another residue. The R group of the acyl fragment may be alkyl, aryl, alkenyl, or alkynyl; the other residue to which the acyl fragment is bonded may be a carbon, hydrogen, oxygen, halogen, nitrogen, sulfur, or other substituent.

**TABLE 1** Some Types of Carbonyl Compounds

Name	General formula	Name ending
Aldehyde	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{H}$	-al
Ketone	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{R}'$	-one
Carboxylic acid	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{H}$	-oic acid
Acid chloride	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{Cl}$	-yl or -oyl chloride
Acid anhydride	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\overset{\text{O}}{\parallel}{\text{C}}-\text{R}'$	-oic anhydride
Ester	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{O}-\text{R}'$	-oate
Lactone (cyclic ester)		None
Amide	$\text{R}-\overset{\text{O}}{\parallel}{\text{C}}-\text{N}$ (with two bonds extending from the nitrogen)	-amide
Lactam (cyclic amide)		None

It's useful to classify carbonyl compounds into two general categories, based on the kinds of chemistry they undergo:

Aldehydes (RCHO)  
Ketones (R<sub>2</sub>CO)

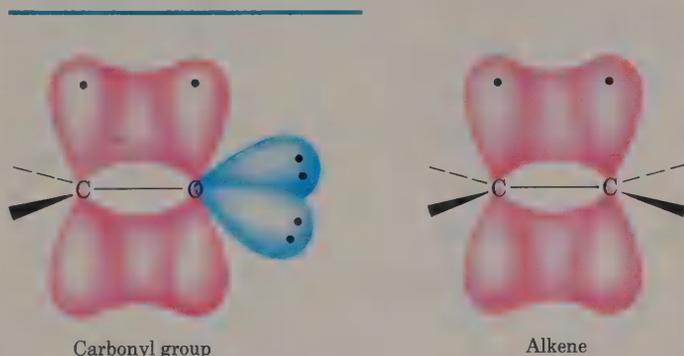
Carboxylic acids (RCOOH)  
Esters (RCOOR')  
Acid chlorides (RCOCl)  
Acid anhydrides (RCOOCOR')  
Amides (RCONH<sub>2</sub>)

The acyl groups in these two families are bonded to substituents (-H and -R, respectively) that *can't stabilize a negative charge and therefore can't act as leaving groups*. Aldehydes and ketones behave similarly and undergo many of the same reactions.

The acyl groups in carboxylic acids and their derivatives are bonded to substituents (oxygen, halogen, nitrogen) that *can stabilize a negative charge and can serve as leaving groups in substitution reactions*. The chemistry of these compounds is therefore similar.

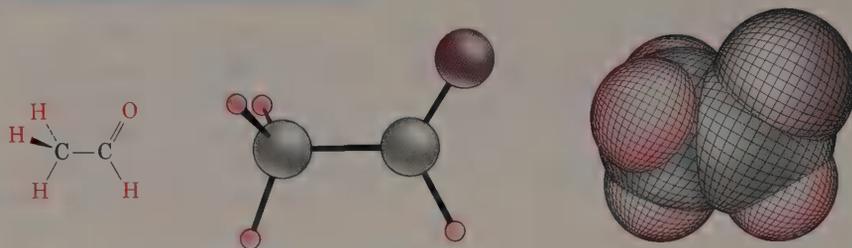
## II. Nature of the Carbonyl Group

The carbon-oxygen double bond of carbonyl groups is similar in many respects to the carbon-carbon double bond of alkenes (Figure 1). The carbonyl carbon atom is  $sp^2$  hybridized and forms three sigma bonds. The fourth valence electron remains in a carbon  $p$  orbital and forms a pi bond to oxygen by overlap with an oxygen  $p$  orbital. The oxygen atom also has two non-bonding pairs of electrons, which occupy its remaining two orbitals.



**Figure 1** Electronic structure of the carbonyl group.

Like alkenes, carbonyl compounds are planar about the double bond and have bond angles of approximately 120°. Figure 2 shows the structure of acetaldehyde and indicates the experimentally determined bond lengths and angles. As you might expect, the carbon-oxygen double bond is both shorter (1.22 Å versus 1.43 Å) and stronger [175 kcal/mol (732 kJ/mol) versus 92 kcal/mol (385 kJ/mol)] than a carbon-oxygen single bond.



<i>Bond angle (°)</i>		<i>Bond length (Å)</i>	
H—C—C	118	C=O	1.22
C—C—O	121	C—C	1.50
H—C—O	121	OC—H	1.09

**Figure 1** Structure of acetaldehyde.

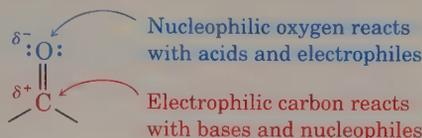
Carbon–oxygen double bonds are polarized because of the high electronegativity of oxygen relative to carbon. Thus, all types of carbonyl compounds have substantial dipole moments, as shown in Table 2.

**TABLE 2** Dipole Moments of Some Carbonyl Compounds,  $R_2CO$

<i>Carbonyl compound</i>	<i>Type of carbonyl group</i>	<i>Observed dipole moment (D)</i>
HCHO	Aldehyde	2.33
CH <sub>3</sub> CHO	Aldehyde	2.72
(CH <sub>3</sub> ) <sub>2</sub> CO	Ketone	2.88
PhCOCH <sub>3</sub>	Ketone	3.02
Cyclobutanone	Ketone	2.99
CH <sub>3</sub> COOH	Carboxylic acid	1.74
CH <sub>3</sub> COCl	Acid chloride	2.72
CH <sub>3</sub> CO <sub>2</sub> CH <sub>3</sub>	Ester	1.72
CH <sub>3</sub> CONH <sub>2</sub>	Amide	3.76
CH <sub>3</sub> CON(CH <sub>3</sub> ) <sub>2</sub>	Amide	3.81

The most important consequence of carbonyl-group polarization is the chemical reactivity of the carbon–oxygen double bond. Since the carbonyl carbon carries a partial positive charge, it is an electrophilic site and is attacked by nucleophiles. Conversely, the carbonyl oxygen carries a partial

negative charge and is a nucleophilic (basic) site. We'll see in the next five chapters that the majority of carbonyl-group reactions can be rationalized in terms of simple bond-polarization arguments.

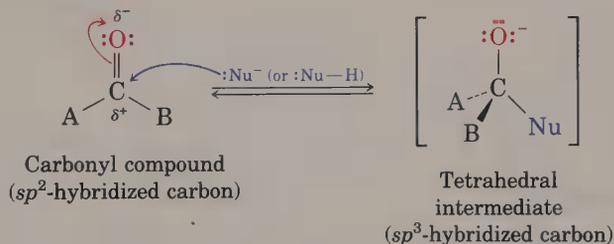


### III. General Reactions of Carbonyl Compounds

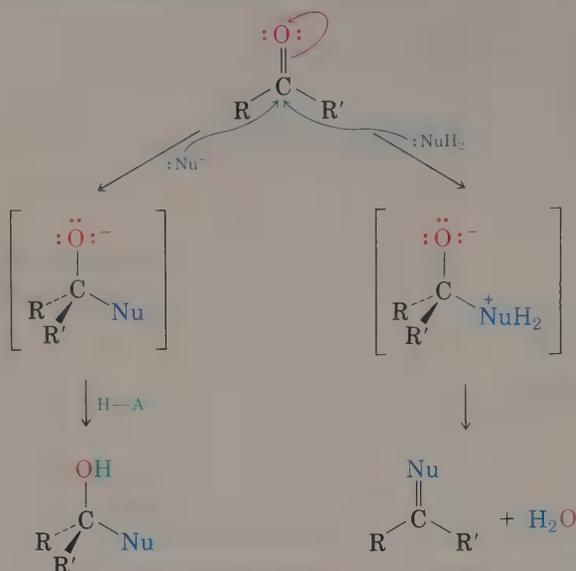
Most reactions of carbonyl groups occur by one of four general mechanisms: **nucleophilic addition**, **nucleophilic acyl substitution**, **alpha substitution**, and **carbonyl condensation**. These mechanisms have many variations, just as alkene electrophilic addition reactions and  $S_N2$  reactions do, but the variations are much easier to learn when the fundamental features of the mechanisms are understood. Let's see what these four reaction mechanisms are and what kinds of chemistry carbonyl groups undergo.

#### NUCLEOPHILIC ADDITION REACTIONS

The most common reaction of ketones and aldehydes is the **nucleophilic addition reaction**, in which a nucleophile adds to the electrophilic carbon of the carbonyl group. Since the nucleophile uses its electron pair to form a new bond to carbon, two electrons from the carbon–oxygen double bond must move onto the electronegative oxygen atom, where they can be stabilized in an alkoxide anion. The carbonyl carbon rehybridizes from  $sp^2$  to  $sp^3$  during the reaction, and the initial intermediate therefore has tetrahedral geometry.



Once formed, and depending on the nature of the nucleophile, the tetrahedral intermediate can undergo either of the reactions shown in Figure 3. Often, the tetrahedral alkoxide intermediate is simply protonated by water or acid to form an alcohol product, as happens during Grignard addition (Section 17.7). Alternatively, the tetrahedral intermediate can expel the oxygen completely to form a new double bond between the carbonyl-group

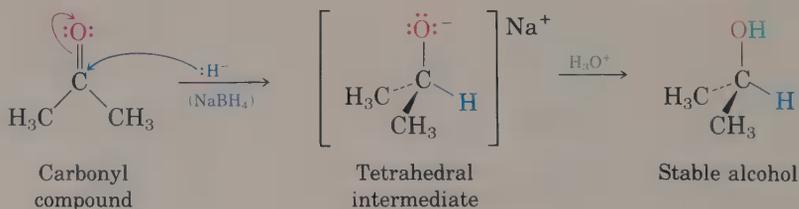


**Figure 3** The addition reaction of a nucleophile with a ketone or an aldehyde.

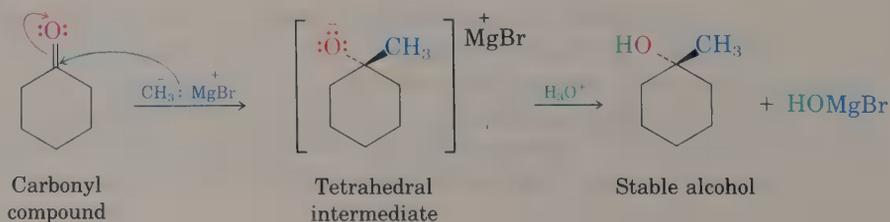
carbon and the nucleophile. This second reaction is particularly common when amine nucleophiles are involved.

**Alcohol Formation** The simplest reaction of a tetrahedral intermediate is protonation to yield an alcohol as final product. We've already seen two examples of this kind of process during reduction of ketones and aldehydes with hydride reagents such as  $\text{NaBH}_4$  and  $\text{LiAlH}_4$  (Section 17.6) and during Grignard reactions (Section 17.7). In the case of reduction, the nucleophile that adds to the carbonyl group is a **hydride ion,  $\text{H}^-$** , while in the case of Grignard reaction, the nucleophile is a **carbanion,  $\text{R}_3\text{C}^-$** .

Hydride reduction

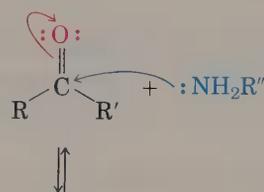


Grignard reaction

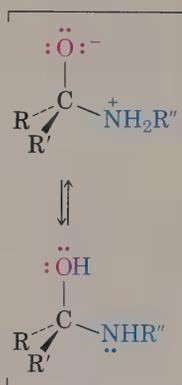


**Formation of C=Nu** The second mode of nucleophilic addition, which often happens with amine nucleophiles, is that complete loss of oxygen occurs and a C=Nu double bond forms. For example, ketones and aldehydes react with primary amines,  $R'NH_2$ , to form **imines**,  $R_2C=NR'$ . These reactions proceed through exactly the same kind of tetrahedral intermediate as that formed during hydride reduction and Grignard reaction, but the initially formed product is not isolated. Instead, it loses water to form an imine, as shown in Figure 4.

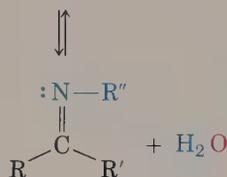
Addition to the ketone or aldehyde carbonyl group by the neutral amine nucleophile gives a dipolar tetrahedral intermediate.



Transfer of a proton from nitrogen to oxygen then yields a nonpolar amino alcohol intermediate.



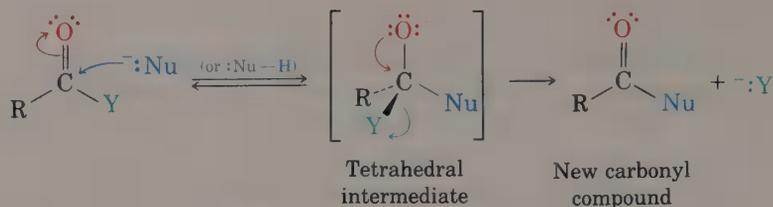
Dehydration of the amino alcohol intermediate gives neutral imine plus water as final products.



**Figure 4** Mechanism of imine formation by reaction of an amine with a ketone or an aldehyde.

## NUCLEOPHILIC ACYL SUBSTITUTION REACTIONS

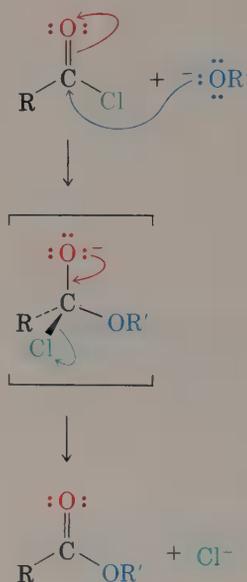
The second fundamental reaction of carbonyl compounds, **nucleophilic acyl substitution**, is related to the nucleophilic addition reaction just discussed but occurs only with carboxylic acid derivatives rather than with ketones and aldehydes. When the carbonyl group of a carboxylic acid derivative reacts with a nucleophile, addition occurs, but the initially formed tetrahedral intermediate is not isolated. Since carboxylic acid derivatives have a leaving group bonded to the carbonyl-group carbon, the tetrahedral intermediate can react further by expelling the leaving group and forming a new carbonyl compound:



where Y = —OR (ester), —Cl (acid chloride), —NH<sub>2</sub> (amide), or —OCOR' (acid anhydride)

The net effect of nucleophilic acyl substitution is the replacement of the leaving group by the attacking nucleophile. We'll see in Chapter 21, for example, that acid chlorides are rapidly converted into esters by treatment with alkoxide ions. This reaction is simply a nucleophilic acyl substitution in which alkoxide replaces chloride (Figure 5).

Nucleophilic addition of alkoxide ion to an acid chloride yields a tetrahedral intermediate.

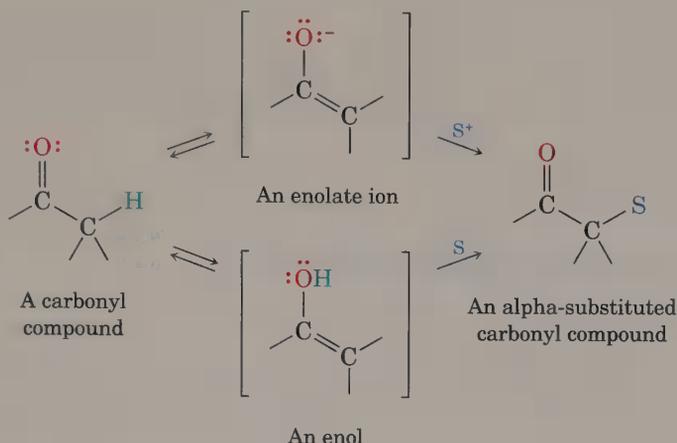


An electron pair from oxygen expels chloride ion and yields the substitution product, an ester.

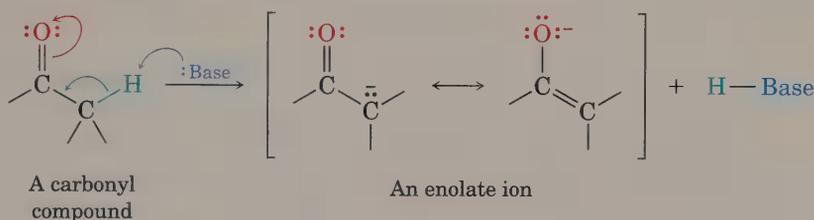
**Figure 5** Mechanism of the nucleophilic acyl substitution reaction of an acid chloride with an alkoxide ion to yield an ester.

## ALPHA-SUBSTITUTION REACTIONS

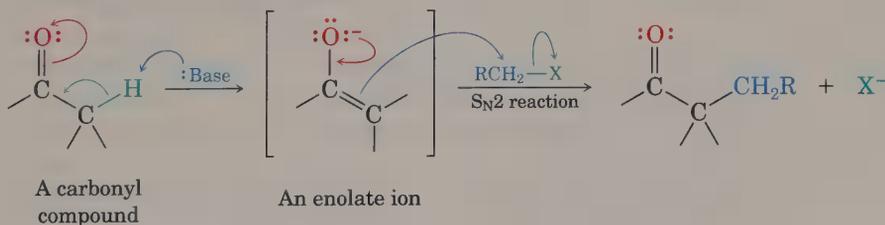
The third major reaction of carbonyl compounds, **alpha substitution**, occurs at the position *next* to the carbonyl group, the alpha ( $\alpha$ ) position. These processes involve the substitution of an alpha hydrogen by some other group and take place by formation of intermediate **enols** or **enolate ions**:



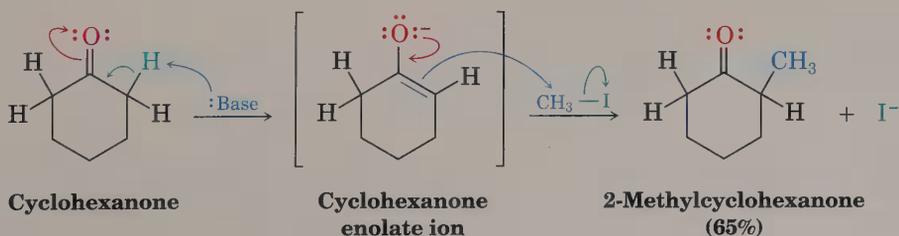
For reasons that we'll explore in Chapter 22, the presence of a carbonyl group renders the protons on the alpha carbon acidic. Carbonyl compounds therefore react with strong base to yield enolate ions.



What chemistry might we expect of enolate ions? Since they carry negative charges and have high electron density on the alpha carbon, enolate ions react as nucleophiles and undergo many of the reactions we've already studied. For example, enolates react with primary alkyl halides in the  $S_N2$  reaction. The nucleophilic carbanion displaces halide ion and a new carbon-carbon bond forms:



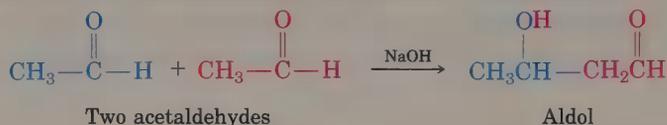
For example:



The  $S_N2$  alkylation reaction between an enolate ion and an alkyl halide is one of the most powerful methods available for making carbon–carbon bonds, thereby building up larger molecules from small precursors. We will study the alkylation of many kinds of carbonyl groups in detail in Chapter 22.

## CARBONYL CONDENSATION REACTIONS

The fourth and last fundamental reaction of carbonyl groups, **carbonyl condensation**, takes place when two carbonyl compounds react with each other. For example, when acetaldehyde is treated with base, two molecules react with each other to yield the hydroxy aldehyde product known as *aldol* (aldehyde + alcohol):



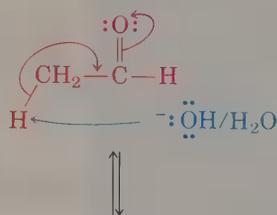
Although the carbonyl condensation reaction appears different from the three general carbonyl-group processes already discussed, it's actually quite similar. A carbonyl condensation reaction is simply a *combination* of a nucleophilic addition step and an alpha-substitution step. The initially formed enolate ion of acetaldehyde acts as a nucleophile and adds to the carbonyl group of another acetaldehyde molecule. Reaction occurs by the pathway shown in Figure 6.

## IV. Summary

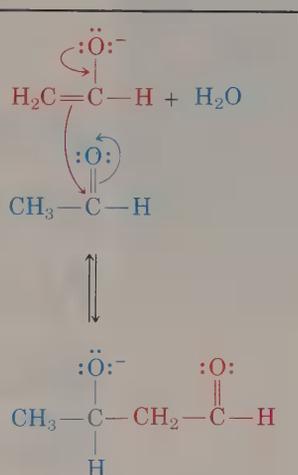
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The purpose of this short overview of carbonyl-group reactions is not to present details of specific reactions but rather to lay the groundwork for the following chapters. Although many of the reactions we'll see in the next five chapters appear unrelated at first glance, almost every reaction we'll encounter is a **nucleophilic addition**, a **nucleophilic acyl substitution**, an **alpha substitution**, or a **carbonyl condensation**. Knowing where we'll be heading should help you to keep matters straight on this most important of all functional groups, the carbonyl group.

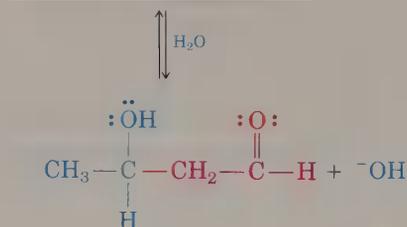
Hydroxide ion abstracts an acidic alpha proton from one molecule of acetaldehyde, yielding an enolate ion.



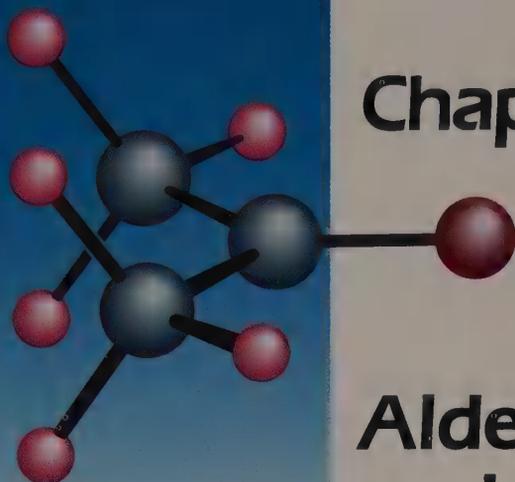
The enolate ion adds as a nucleophile to the carbonyl group of a second molecule of acetaldehyde, producing a tetrahedral intermediate.



The intermediate is protonated by water solvent to yield the neutral aldol product and regenerate hydroxide ion.



**Figure 6** Mechanism of a carbonyl condensation reaction between two molecules of acetaldehyde.

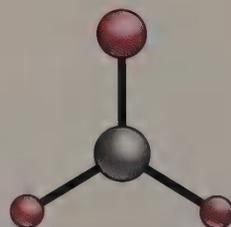
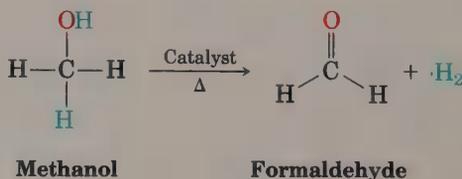


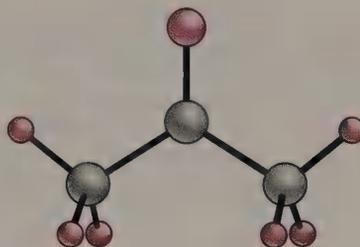
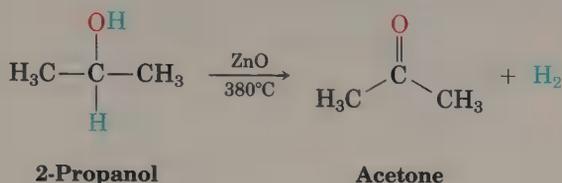
## Chapter 19

# Aldehydes and Ketones: Nucleophilic Addition Reactions

Aldehydes and ketones are among the most important of all compounds, both in nature and in the chemical industry. In nature, many substances required by living systems are aldehydes or ketones. In the chemical industry, simple aldehydes and ketones are synthesized in large quantities for use both as solvents and as starting materials for a host of other products. For example, approximately 3.4 million tons per year of formaldehyde,  $\text{H}_2\text{C}=\text{O}$ , are manufactured for use in building insulation materials and in the adhesive resins that bind particle board and plywood. Concern over the possible toxicity of such materials may sharply curtail their uses, however. Acetone,  $(\text{CH}_3)_2\text{C}=\text{O}$ , is also widely used in industry, with some 1.2 million tons per year prepared for use as a solvent.

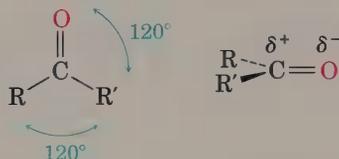
Formaldehyde is synthesized industrially by catalytic dehydrogenation (oxidation) of methanol, and one method of acetone preparation involves dehydrogenation of 2-propanol. These methods are of little use in the laboratory, however, because of the high temperatures required.





## 19.1 Properties of Aldehydes and Ketones

As we saw in the previous overview, the carbonyl functional group is planar, with bond angles of approximately  $120^\circ$ , and the carbon–oxygen double bond is polar.



One consequence of carbonyl bond polarity is that aldehydes and ketones are weakly associated and therefore have higher boiling points than alkanes of similar molecular weight have. Since ketones and aldehydes can't form hydrogen bonds, however, their boiling points are lower than those of the corresponding alcohols. Formaldehyde, the simplest aldehyde, is a gas at room temperature, but other simple aldehydes and ketones are liquid (Table 19.1).

## 19.2 Nomenclature of Aldehydes and Ketones

### ALDEHYDES

Aldehydes are named by replacing the terminal *-e* of the corresponding alkane name with *-al*. The longest chain selected as the base name must contain the  $-\text{CHO}$  group, and the  $-\text{CHO}$  carbon is always numbered as carbon 1. For example:

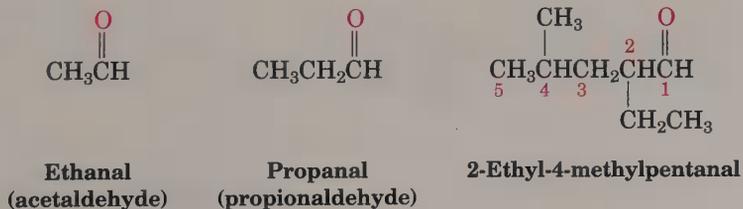
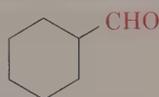


TABLE 19.1 Physical Properties of Simple Aldehydes and Ketones

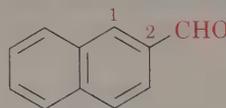
Compound name	Structure	Boiling point (°C)	Melting point (°C)
Formaldehyde	HCHO	-21	-92
Acetaldehyde	CH <sub>3</sub> CHO	21	-121
Propanal	CH <sub>3</sub> CH <sub>2</sub> CHO	49	-81
Butanal	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>2</sub> CHO	76	-99
Pentanal	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>3</sub> CHO	103	-92
Benzaldehyde	C <sub>6</sub> H <sub>5</sub> CHO	178	-26
Acetone	CH <sub>3</sub> COCH <sub>3</sub>	56	-95
2-Butanone	CH <sub>3</sub> CH <sub>2</sub> COCH <sub>3</sub>	80	-86
2-Pentanone	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> COCH <sub>3</sub>	102	-78
3-Pentanone	CH <sub>3</sub> CH <sub>2</sub> COCH <sub>2</sub> CH <sub>3</sub>	102	-40
Cyclohexanone		156	-16

Note that the longest chain in 2-ethyl-4-methylpentanal is a hexane but that this chain does not include the -CHO group and thus is not considered the parent.

For more complex aldehydes in which the -CHO group is attached to a ring, the suffix *-carbaldehyde* is used:



Cyclohexane carbaldehyde

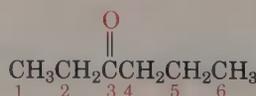
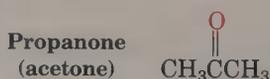


2-Naphthalene carbaldehyde

Certain simple and well-known aldehydes have common names that are recognized by IUPAC. Some of the more important common names are given in Table 19.2.

## KETONES

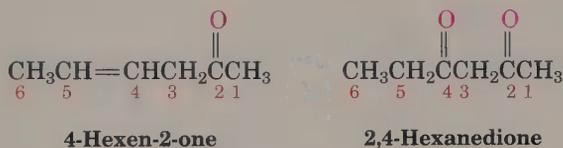
Ketones are named by replacing the terminal *-e* of the corresponding alkane name with *-one* (pronounced own). The chain selected for the base name is the longest one that contains the ketone group, and the numbering begins at the end nearer the carbonyl carbon. For example:



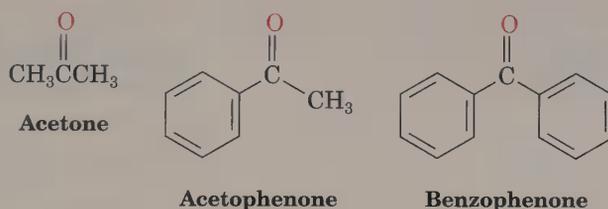
3-Hexanone

TABLE 19.2 Common Names of Some Simple Aldehydes

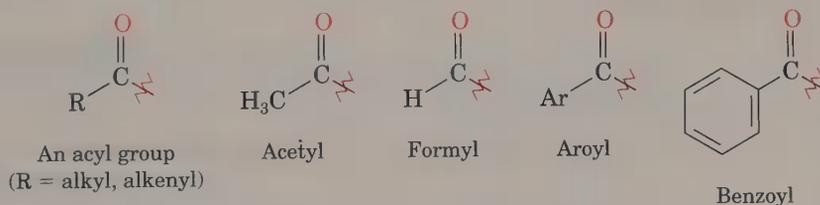
Formula	Common name	Systematic name
HCHO	Formaldehyde	Methanal
CH <sub>3</sub> CHO	Acetaldehyde	Ethanal
CH <sub>3</sub> CH <sub>2</sub> CHO	Propionaldehyde	Propanal
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CHO	Butyraldehyde	Butanal
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CHO	Valeraldehyde	Pentanal
H <sub>2</sub> C=CHCHO	Acrolein	2-Propenal
 CHO	Benzaldehyde	Benzenecarbaldehyde



Certain ketones are allowed by IUPAC to retain their common names, though these are few:



When it becomes necessary to refer to the RCO- group as a substituent, the term **acyl** (a-sil) is used and the name ending *-yl* is attached. For example, the CH<sub>3</sub>CO- group is an **acetyl** group. Similarly, -CHO is called a **formyl** group, and ArCO- is referred to as an **aroyl** group.

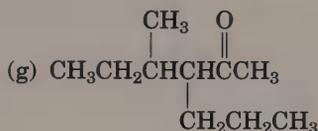
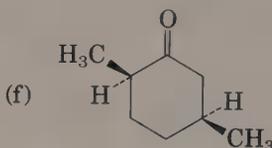
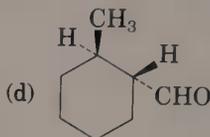
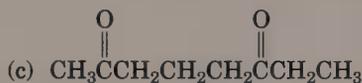
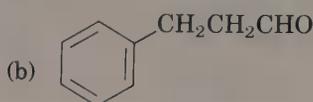
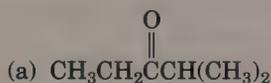


If other functional groups are present and the doubly bonded oxygen must be considered a substituent, the prefix *oxo-* is used. For example:



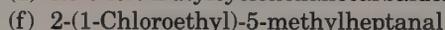
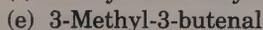
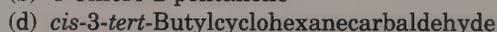
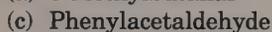
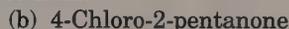
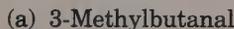
## PROBLEM.....

19.1 Name these aldehydes and ketones according to IUPAC rules:



## PROBLEM.....

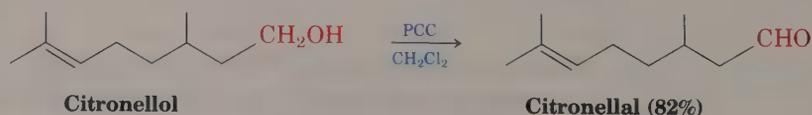
19.2 Draw structures corresponding to these names:



## 19.3 Preparation of Aldehydes

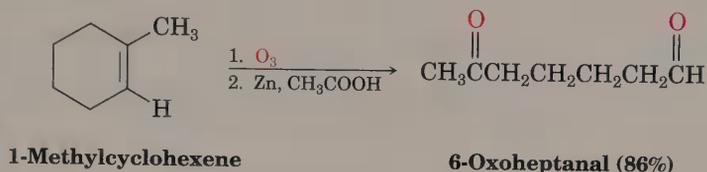
We've already discussed two of the best methods of aldehyde synthesis: oxidation of primary alcohols and oxidative cleavage of alkenes. Let's review them briefly.

1. Primary alcohols can be oxidized to give aldehydes (Section 17.9). The reaction is often carried out using pyridinium chlorochromate (PCC) in dichloromethane solvent at room temperature, and high yields are usually obtained.

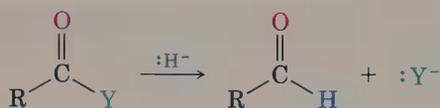


2. Alkenes with at least one vinylic hydrogen undergo oxidative cleavage when treated with ozone to yield aldehydes (Section 7.9). If the

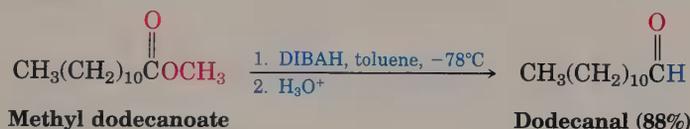
ozonolysis reaction is carried out on cyclic alkenes, dicarbonyl compounds result.



Yet a third method of aldehyde synthesis is one that we'll mention here just briefly and then return to for a more detailed explanation in Section 21.5. Certain carboxylic acid derivatives can be partially reduced to yield aldehydes.



For example, the partial reduction of esters by diisobutylaluminum hydride (DIBAH) is an extremely important laboratory-scale method of aldehyde synthesis. The reaction is normally carried out at  $-78^\circ\text{C}$  (dry-ice temperature) in toluene solution, and yields are often excellent.



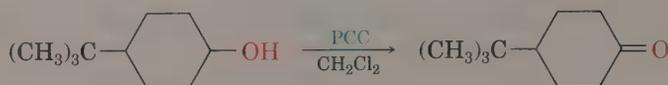
**PROBLEM** .....

- 19.3 How would you prepare pentanal from these starting materials?  
 (a) 1-Pentanol (b) 1-Hexene (c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{COOCH}_3$
- .....

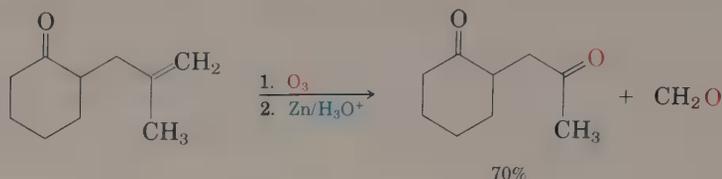
## 19.4 Preparation of Ketones

For the most part, methods of ketone synthesis are analogous to those for aldehydes:

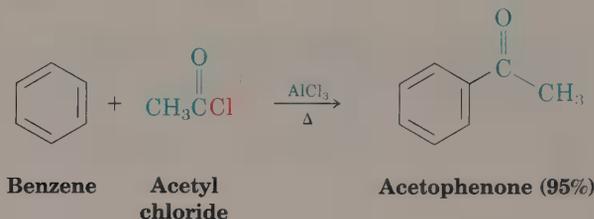
1. Secondary alcohols are oxidized by a variety of reagents to give ketones (Section 17.9). Jones' reagent ( $\text{CrO}_3$  in aqueous sulfuric acid), pyridinium chlorochromate, and sodium dichromate in aqueous acetic acid are all effective, and the specific choice of reagent depends on such factors as reaction scale, cost, and acid or base sensitivity of the alcohol.

4-*tert*-Butylcyclohexanol4-*tert*-Butylcyclohexanone (90%)

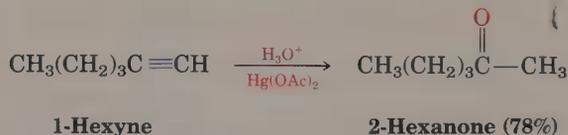
2. Ozonolysis of alkenes yields ketones if one of the unsaturated carbon atoms is disubstituted (Section 7.9):



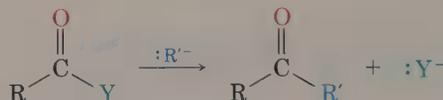
3. Aryl ketones are prepared by Friedel–Crafts acylation of an aromatic ring with an acid chloride in the presence of  $\text{AlCl}_3$  catalyst (Section 16.4):



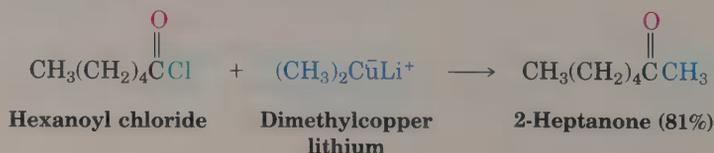
4. Methyl ketones are prepared by hydration of terminal alkynes in the presence of mercuric ion as catalyst (Section 8.5). Internal alkynes can also be hydrated, but a mixture of ketone products usually results.



5. Ketones can be prepared from certain carboxylic acid derivatives, just as aldehydes can. We'll discuss this subject in more detail in Section 21.5.



Among the most useful syntheses of this type is the reaction between an acid chloride and a diorganocopper reagent. This process is reminiscent of the coupling reaction between alkyl halides and diorganocoppers that we saw in Section 10.10:



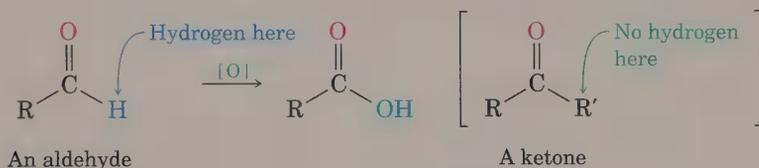
## PROBLEM.....

19.4 How would you carry out these reactions? More than one step may be required.

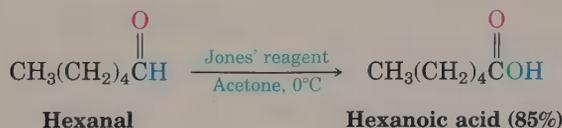
- (a) 3-Hexyne  $\longrightarrow$  3-Hexanone  
 (b) Benzene  $\longrightarrow$  *m*-Bromoacetophenone  
 (c) Bromobenzene  $\longrightarrow$  Acetophenone  
 (d) 1-Methylcyclohexene  $\longrightarrow$  2-Methylcyclohexanone

## 19.5 Oxidation of Aldehydes and Ketones

Aldehydes are readily oxidized to yield carboxylic acids, but ketones are unreactive toward oxidation except under the most vigorous conditions. This difference in behavior is a consequence of the structural difference between the two functional groups: Aldehydes have a  $-\text{CHO}$  proton that can be readily abstracted during oxidation, but ketones do not.

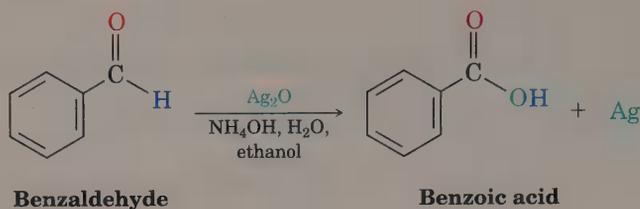


Many oxidizing agents such as hot nitric acid and potassium permanganate convert aldehydes into carboxylic acids, but the Jones reagent,  $\text{CrO}_3$  in aqueous sulfuric acid, is a more common choice on a small laboratory scale (Section 17.9). Jones oxidations occur rapidly at room temperature and give good yields of product.



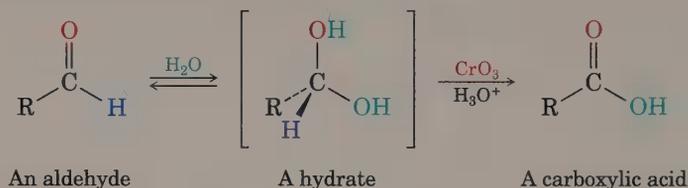
One drawback to the Jones oxidation is that the conditions under which it occurs are acidic, and sensitive molecules sometimes undergo acid-catalyzed decomposition. Where such decomposition occurs, aldehyde oxidations are often carried out using a dilute ammonia solution of silver oxide,  $\text{Ag}_2\text{O}$  (the **Tollens<sup>1</sup> reagent**). Aldehydes are oxidized by the Tollens reagent in high yield without harming carbon-carbon double bonds or other functional groups in the molecule.

<sup>1</sup>Bernhard Tollens (1841–1918); b. Hamburg, Germany; Ph.D. Göttingen; professor, University of Göttingen.

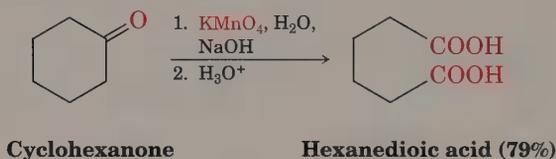


A shiny mirror of metallic silver is deposited on the walls of the flask during a Tollens oxidation, forming the basis of a qualitative test for the presence of an aldehyde functional group in a molecule of unknown structure. A small sample of the unknown is dissolved in ethanol in a test tube, and a few drops of the Tollens reagent are added. If the test tube becomes silvery, the unknown is presumed to be an aldehyde.

Aldehyde oxidations occur through intermediate 1,1-diols or hydrates, which are formed by a reversible nucleophilic addition of water to the carbonyl group. Even though present in small amounts at equilibrium, the hydrate reacts like any normal primary or secondary alcohol and is oxidized to a carbonyl compound (Section 17.9).



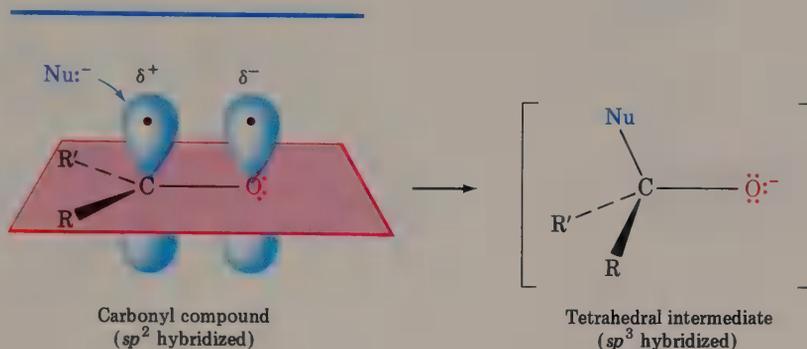
Ketones are inert to most common oxidizing agents but undergo a slow cleavage reaction when treated with hot alkaline  $\text{KMnO}_4$ . The carbon-carbon bond next to the carbonyl group is broken and carboxylic acid fragments are produced. The reaction is useful only for symmetrical ketones such as cyclohexanone, however, since product mixtures are formed from unsymmetrical ketones.



## 19.6 Nucleophilic Addition Reactions of Aldehydes and Ketones

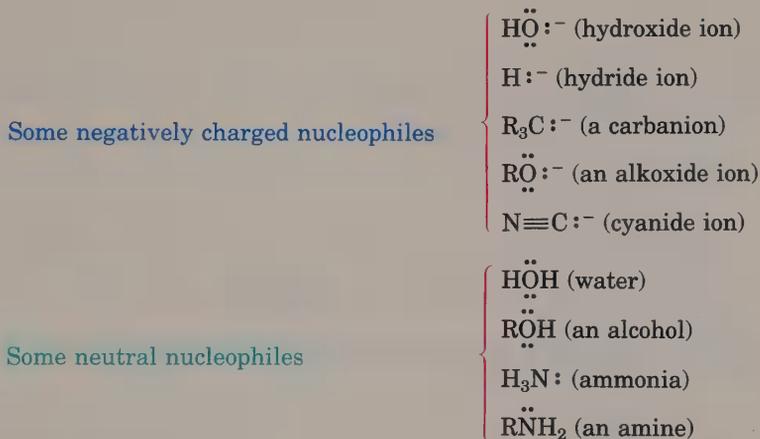
The most important reaction of ketones and aldehydes is the **nucleophilic addition reaction**. A nucleophile attacks the electrophilic carbon atom of the polar carbonyl group from a direction approximately perpendicular to the plane of the carbonyl  $sp^2$  orbitals. Rehybridization of the carbonyl carbon

from  $sp^2$  to  $sp^3$  then occurs, and a tetrahedral alkoxide ion intermediate is produced (Figure 19.1).



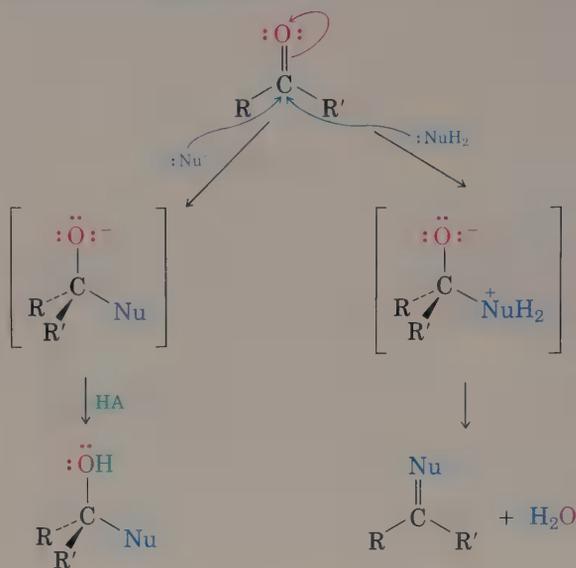
**Figure 19.1** A nucleophilic addition reaction to a ketone or an aldehyde.

The attacking nucleophile can be either negatively charged ( $\text{Nu}^-$ ) or neutral ( $:\text{Nu}-\text{H}$ ). If it's neutral, however, the nucleophile is usually attached to a hydrogen atom that can subsequently be eliminated. For example:



Nucleophilic additions to ketones and aldehydes have two variations, as shown in Figure 19.2: (1) The tetrahedral intermediate can be protonated by water or acid to give an alcohol, or (2) the carbonyl oxygen atom can be expelled as  $\text{HO}^-$  or  $\text{H}_2\text{O}$  to give a new carbon-nucleophile double bond.

In the remainder of this chapter, we'll examine some specific examples of nucleophilic addition reactions. In looking at the details of each reaction, we'll be concerned with two key points: the *reversibility* of a given reaction and the acid or base *catalysis* of that reaction. Some nucleophilic addition reactions take place without catalysis, but many others require acid or base to proceed.



**Figure 19.2** Two possible reaction pathways following addition of a nucleophile to a ketone or an aldehyde.

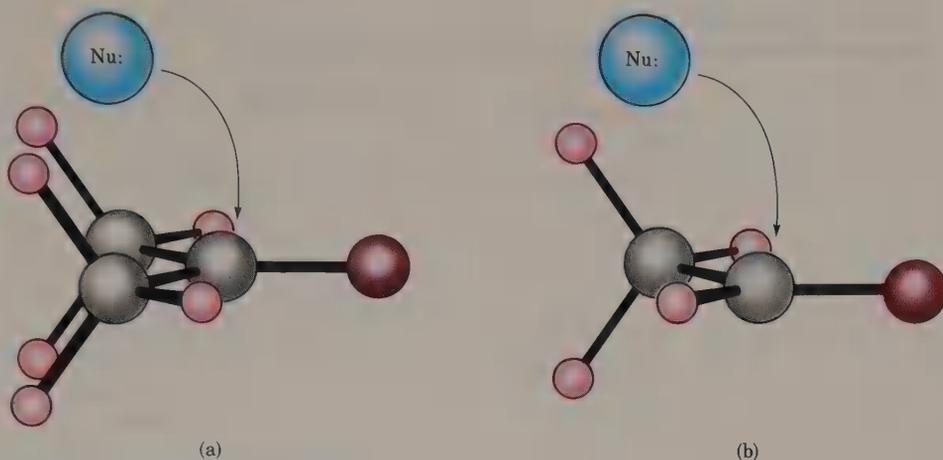
**PROBLEM** .....

- 19.5** What product would you obtain if cyanide ion ( $\text{:C}\equiv\text{N}^-$ ) were added to acetone and the tetrahedral intermediate were protonated?
- .....

## 19.7 Relative Reactivity of Aldehydes and Ketones

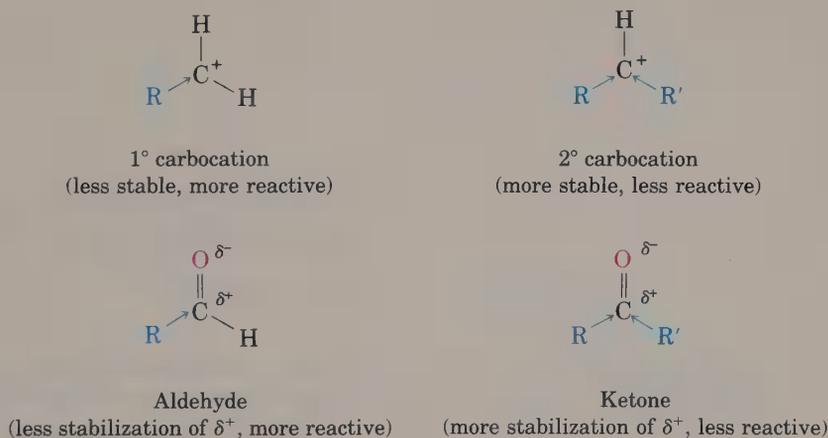
Aldehydes are generally more reactive than ketones in nucleophilic addition reactions for both steric and electronic reasons. Sterically, the presence of two relatively large substituents in ketones versus only one large substituent in aldehydes means that attacking nucleophiles are able to approach aldehydes more readily. Thus, the transition state leading to the tetrahedral intermediate is less crowded and lower in energy for aldehydes than for ketones (Figure 19.3).

Electronically, aldehydes are more reactive than ketones because of the greater degree of polarity of aldehyde carbonyl groups. The easiest way to see this polarity difference is to recall the stability order of carbocations (Section 6.11). Primary carbocations are less stable than secondary ones because there is only one alkyl group inductively stabilizing the positive charge, rather than two. For similar reasons, aldehydes are less stable (and therefore more reactive) than ketones because there is only one alkyl group



**Figure 19.3** Nucleophilic attack on a ketone (a) is sterically hindered because of the two relatively large substituents attached to the carbonyl-group carbon. An aldehyde (b) has only one large substituent and is less hindered.

inductively stabilizing the partial positive charge on the carbonyl carbon, rather than two.



PROBLEM .....

**19.6** Aromatic aldehydes such as benzaldehyde are less reactive in nucleophilic addition reactions than aliphatic aldehydes are. Explain.

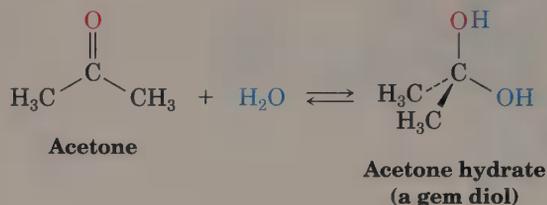
PROBLEM .....

**19.7** Which would you expect to be more reactive toward nucleophilic additions: *p*-methoxybenzaldehyde or *p*-nitrobenzaldehyde? Explain.

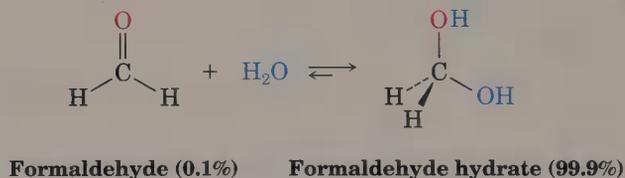
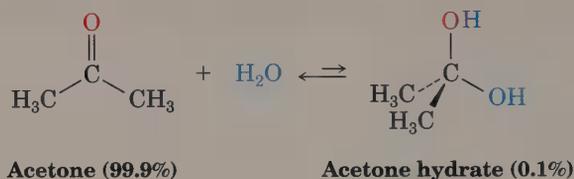
.....

## 19.8 Nucleophilic Addition of H<sub>2</sub>O: Hydration

Aldehydes and ketones undergo reaction with water to yield 1,1-diols, or **geminal (gem) diols**. The hydration reaction is reversible, and gem diols can eliminate water to regenerate ketones or aldehydes.



The exact position of the equilibrium between gem diols and ketones or aldehydes depends on the structure of the carbonyl compound. Although the equilibrium strongly favors the carbonyl compound in most cases, the gem diol is favored for a few simple aldehydes. For example, an aqueous solution of acetone consists of about 0.1% gem diol and 99.9% ketone, whereas an aqueous solution of formaldehyde consists of 99.9% gem diol and 0.1% aldehyde.



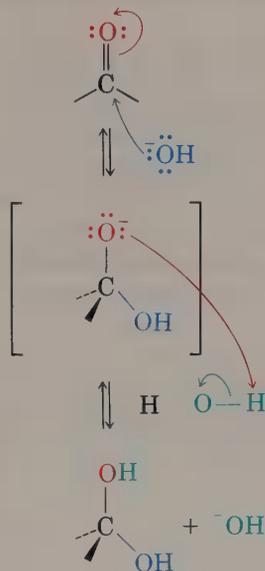
The nucleophilic addition of water to ketones and aldehydes is slow in pure water but is catalyzed by both acid and base. Like all catalysts, acid and base don't change the *position* of the equilibrium; they affect only the rate with which the hydration reaction occurs.

The base-catalyzed reaction takes place in several steps, as shown in Figure 19.4. The attacking nucleophile is the negatively charged hydroxide ion, rather than neutral water.

The acid-catalyzed hydration reaction also takes place in several steps. Acid catalyst first protonates the Lewis-basic oxygen atom of the carbonyl group, placing a positive charge on oxygen and thus making the carbonyl group more electrophilic. Subsequent nucleophilic addition of water to the protonated ketone or aldehyde then occurs to yield a protonated gem diol. Loss of a proton from oxygen then gives the neutral gem diol product (Figure 19.5).

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Hydroxide ion nucleophile adds to the ketone or aldehyde carbonyl group to yield an alkoxide ion intermediate.



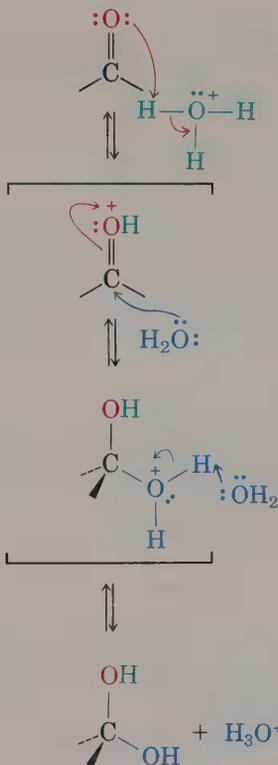
The basic alkoxide ion intermediate abstracts a proton ( $\text{H}^+$ ) from water to yield gem diol product and regenerate hydroxide ion catalyst.

---

**Figure 19.4** Mechanism of base-catalyzed hydration of a ketone or an aldehyde. The purpose of the base is to act as a more reactive nucleophile than neutral water.

---

Acid catalyst protonates the basic carbonyl oxygen atom, making the ketone or aldehyde a much better acceptor of nucleophiles.



Nucleophilic addition of neutral water yields a protonated gem diol.

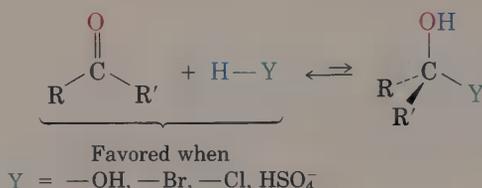
Loss of a proton regenerates the acid catalyst and gives neutral gem diol product.

---

**Figure 19.5** Mechanism of acid-catalyzed hydration of a ketone or an aldehyde. The purpose of the acid catalyst is to protonate the carbonyl starting material, thus making it more electrophilic and more reactive.

Note the important difference between the acid-catalyzed and base-catalyzed processes. The base-catalyzed reaction takes place rapidly because water is converted into hydroxide ion, a much better nucleophilic *donor*. The acid-catalyzed reaction takes place rapidly because the carbonyl compound is converted by protonation into a much better electrophilic *acceptor*.

The hydration reaction just described is typical of what happens when a ketone or an aldehyde is treated with a nucleophile of the type H–Y, where the Y atom is electronegative and can stabilize a negative charge (oxygen, halogen, or sulfur, for example). In such reactions, nucleophilic addition is reversible, with the equilibrium favoring the carbonyl starting material rather than the tetrahedral adduct. Thus, treatment of ketones and aldehydes with reagents such as H<sub>2</sub>O, HCl, HBr, or H<sub>2</sub>SO<sub>4</sub> does not normally lead to addition products.



PROBLEM .....

- 19.8 When dissolved in water, trichloroacetaldehyde (chloral, CCl<sub>3</sub>CHO) exists primarily as the gem diol, chloral hydrate, CCl<sub>3</sub>CH(OH)<sub>2</sub>, better known by the non-IUPAC name “knockout drops.” Show the structure of chloral hydrate.

PROBLEM .....

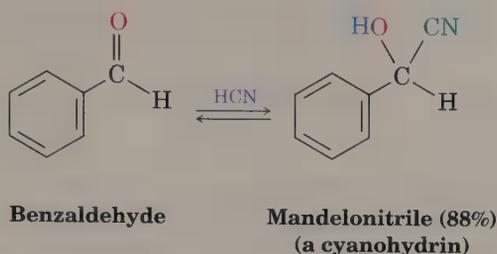
- 19.9 The oxygen in water is primarily (99.8%) <sup>16</sup>O, but water enriched with the heavy isotope <sup>18</sup>O is also available. When a ketone or an aldehyde is dissolved in <sup>18</sup>O-enriched water, the isotopic label becomes incorporated into the carbonyl group: R<sub>2</sub>C=O + H<sub>2</sub>O\* → R<sub>2</sub>C=O\* + H<sub>2</sub>O, where O\* = <sup>18</sup>O. Explain.

PROBLEM .....

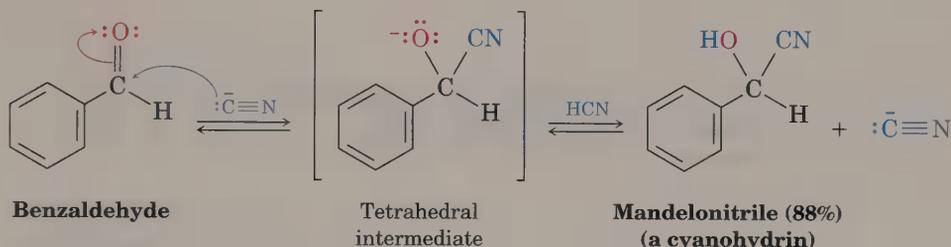
- 19.10 How can you explain the observation that S<sub>N</sub>2 reaction of (dibromomethyl)benzene, C<sub>6</sub>H<sub>5</sub>CHBr<sub>2</sub>, with NaOH yields benzaldehyde rather than (dihydroxymethyl)benzene, C<sub>6</sub>H<sub>5</sub>CH(OH)<sub>2</sub>?
- .....

## 19.9 Nucleophilic Addition of HCN: Cyanohydrins

Aldehydes and unhindered ketones react with HCN to yield **cyanohydrins**, RCH(OH)C≡N. For example, benzaldehyde gives the cyanohydrin mandelonitrile in 88% yield on treatment with HCN:



Detailed studies carried out by Arthur Lapworth<sup>2</sup> in the early 1900s showed that cyanohydrin formation is reversible and base catalyzed. Reaction occurs very slowly when pure HCN is used, but rapidly when a trace amount of base or cyanide ion is added. We can understand this result by recalling that HCN, a weak acid with  $pK_a = 9.1$ , is neither dissociated nor nucleophilic. Cyanide ion, however, is strongly nucleophilic, and addition to ketones and aldehydes occurs by a typical nucleophilic addition pathway. Protonation of the anionic tetrahedral intermediate yields the tetrahedral cyanohydrin product plus regenerated cyanide ion:

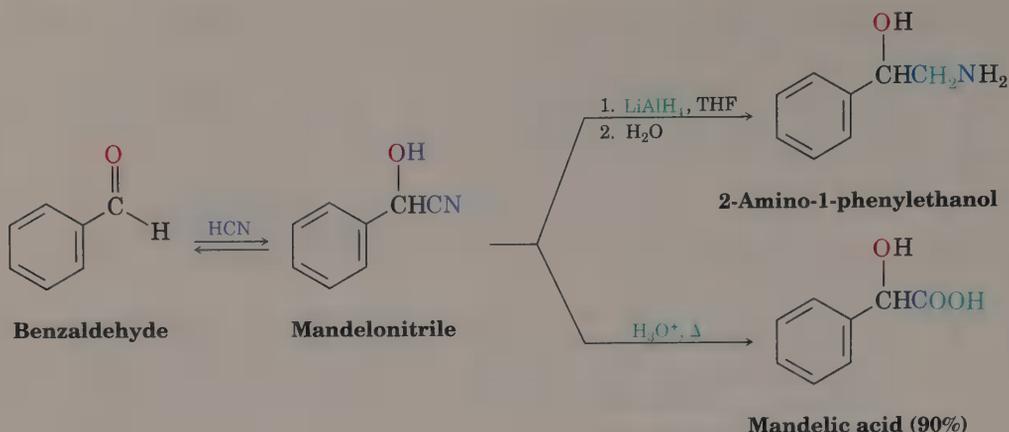


In order to avoid the dangers inherent in handling such a toxic gas as hydrogen cyanide, HCN is usually generated during the reaction by adding a mineral acid like  $H_2SO_4$  to a mixture of carbonyl compound and excess sodium cyanide.

Cyanohydrin formation is particularly interesting because it is one of the few examples of the addition of an acid to a carbonyl group. As noted earlier, acids such as HBr, HCl,  $H_2SO_4$ , and  $CH_3COOH$  don't form carbonyl adducts because the equilibrium constant for reaction is unfavorable. With HCN, however, the equilibrium lies in favor of the adduct.

Cyanohydrin formation is useful because of the further chemistry that can be carried out. For example, nitriles (RCN) can be reduced with  $LiAlH_4$  to yield primary amines ( $RCH_2NH_2$ ) and can be hydrolyzed by aqueous acid to yield carboxylic acids. Thus, cyanohydrin formation provides a method for transforming a ketone or aldehyde into a different functional group while lengthening the carbon chain by one unit.

<sup>2</sup>Arthur Lapworth (1872–1941); b. Galashiels, Scotland; D.Sc., City and Guilds Institute, London; professor, University of Manchester (1909–1941).



PROBLEM .....

- 19.11 How can you account for the observation that cyclohexanone forms a cyanohydrin in good yield but that 2,2,6-trimethylcyclohexanone appears to be unreactive to HCN/KCN?
- .....

## 19.10 Nucleophilic Addition of Grignard Reagents: Alcohol Formation

Although we didn't point it out earlier, the synthesis of alcohols by reaction of Grignard reagents with ketones and aldehydes (Section 17.7) is simply a nucleophilic addition reaction. Unlike the nucleophilic additions of water and HCN, though, Grignard additions are generally irreversible.

Grignard reagents are nucleophiles because the carbon–magnesium bond is strongly polarized, with a high electron density on carbon. Thus, Grignard reagents react as if they were carbanions,  $\text{R}_3\text{C}^-$ . Acid–base complexation of magnesium ion with the carbonyl oxygen atom first serves to make the carbonyl group a better electrophile, and nucleophilic addition then produces a tetrahedrally hybridized magnesium alkoxide intermediate. Protonation of this intermediate by addition of aqueous acid yields the neutral alcohol (Figure 19.6).

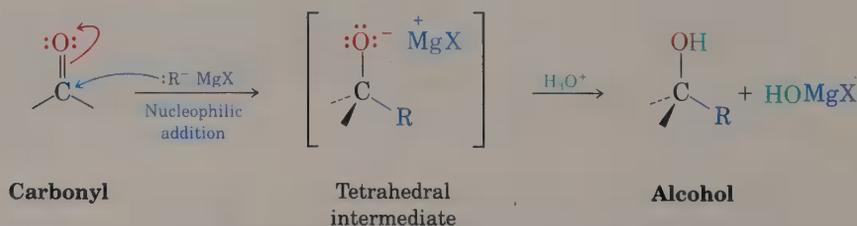
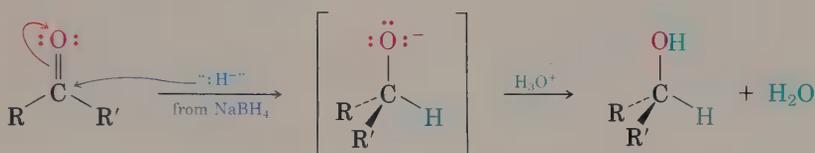


Figure 19.6 Mechanism of the Grignard reaction.

## 19.11 Nucleophilic Addition of Hydride: Reduction

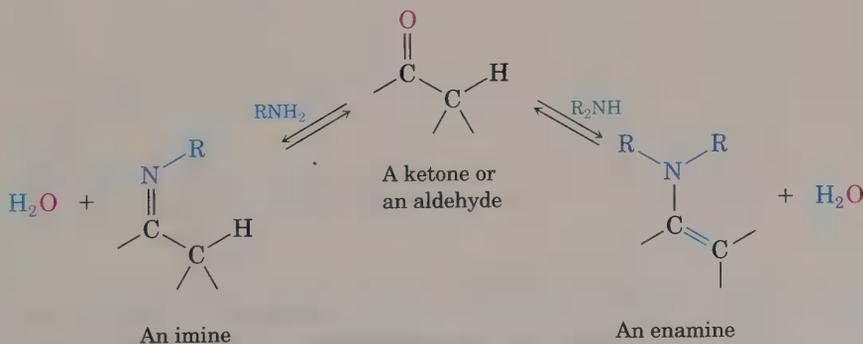
The reduction of ketones and aldehydes to yield alcohols (Section 17.6) is another nucleophilic addition reaction whose mechanism we didn't point out earlier. The exact details of carbonyl-group reduction by hydride agents such as  $\text{LiAlH}_4$  and  $\text{NaBH}_4$  are complex because there is no such species as a discrete hydride ion,  $\text{H}^-$ . Nevertheless, the common reducing agents function as if they were "hydride ion equivalents," and the fundamental step in carbonyl-group reduction is a nucleophilic addition (Figure 19.7). Addition of water or aqueous acid after the hydride addition step protonates the tetrahedral alkoxide intermediate and gives the alcohol product.



**Figure 19.7** Mechanism of carbonyl-group reduction by nucleophilic addition of "hydride ion" from  $\text{NaBH}_4$  or  $\text{LiAlH}_4$ .

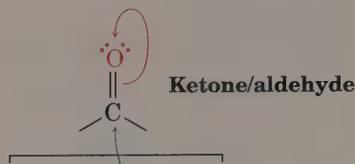
## 19.12 Nucleophilic Addition of Amines: Imine and Enamine Formation

Primary amines,  $\text{RNH}_2$ , add to aldehydes and ketones to yield **imines**,  $\text{R}_2\text{C}=\text{NR}$ . Secondary amines add similarly to yield **enamines** (*-ene* + *amine*; unsaturated amine).

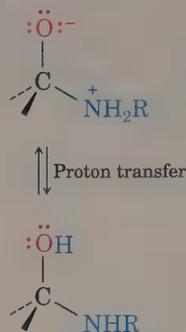


These two reactions appear different, since one leads to a product with a carbon–nitrogen double bond and the other leads to a product with a

Nucleophilic attack on the ketone or aldehyde by the lone-pair electrons of an amine leads to a dipolar tetrahedral intermediate.

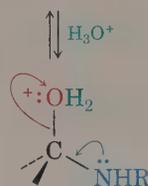


A proton is then transferred from nitrogen to oxygen, yielding a neutral carbinolamine.

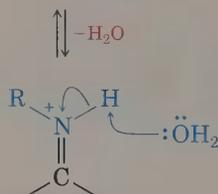


Acid catalyst protonates the hydroxyl oxygen.

**Carbinolamine**

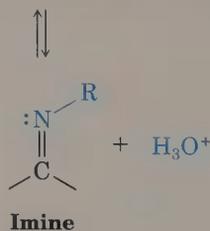


The nitrogen lone-pair electrons expel water, giving an iminium ion.



**Iminium ion**

Loss of  $\text{H}^+$  from nitrogen then gives the neutral imine product.



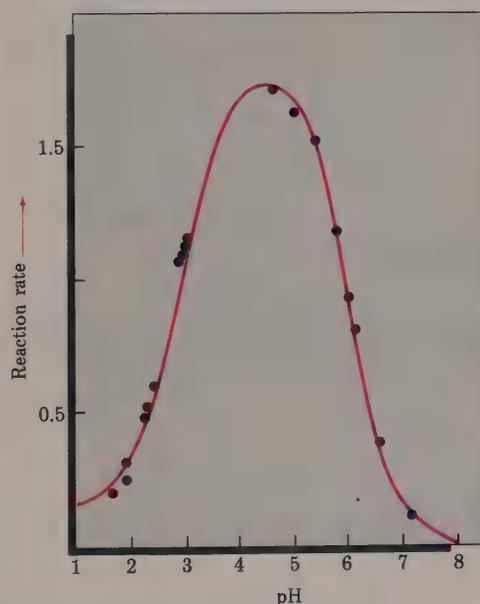
**Figure 19.8** Mechanism of imine formation by reaction of a ketone or an aldehyde with a primary amine.

carbon-carbon double bond, but they're actually quite similar. Both are typical examples of nucleophilic addition reactions in which the initially formed tetrahedral intermediate can't be isolated. Instead, water is eliminated, and a new carbon-nucleophile double bond is formed.

Imines are formed by a reversible, acid-catalyzed process involving nucleophilic attack on the carbonyl group by the primary amine, followed by transfer of a proton from nitrogen to oxygen to yield a neutral amino alcohol (a **carbinolamine**). Protonation of the carbinolamine oxygen by the acid catalyst then converts the hydroxyl into a better leaving group, and E1-like loss of water produces an iminium ion. Loss of a proton then gives the final product and regenerates acid catalyst (Figure 19.8 on page 724).

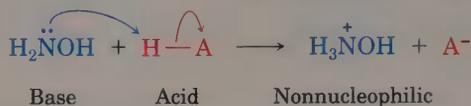
Imine formation is normally an acid-catalyzed process. Studies have revealed a pH versus rate profile indicating that reaction is very slow at both high and low pH but reaches a maximum rate at weakly acidic pH. For example, Figure 19.9 shows the profile obtained for reaction between acetone and hydroxylamine,  $\text{H}_2\text{NOH}$ , indicating that maximum reaction rate is obtained at pH 4.5.

We can explain the observed maximum at pH 4.5 by looking at each individual step in the mechanism. As indicated in Figure 19.8, acid is required to protonate the intermediate carbinolamine, thereby converting the hydroxyl into a better leaving group. Thus, reaction can't occur if there is not enough acid present (high pH). On the other hand, if too much acid



**Figure 19.9** Dependence on pH of the rate of reaction between acetone and hydroxylamine,  $\text{H}_2\text{NOH}$ :  $(\text{CH}_3)_2\text{C}=\text{O} + \text{H}_2\text{NOH} \rightarrow (\text{CH}_3)_2\text{C}=\text{NOH} + \text{H}_2\text{O}$ .

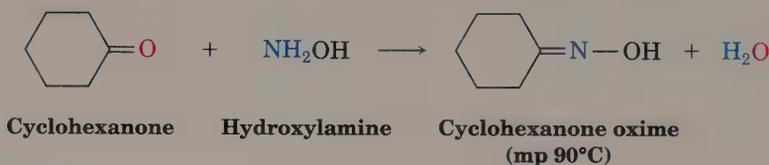
is present (low pH), the attacking amine nucleophile is completely protonated and the initial nucleophilic addition step can't occur.



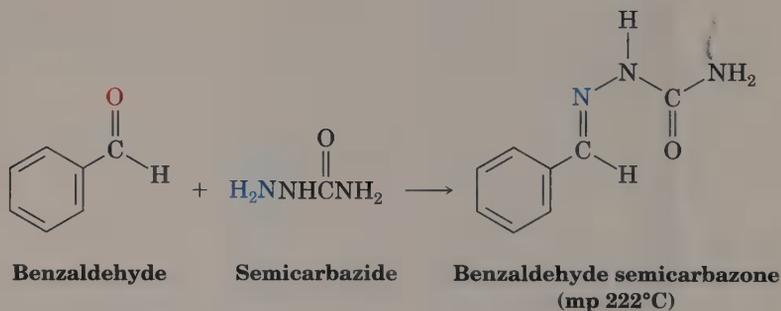
Evidently, pH 4.5 represents a compromise between the need for *some* acid to catalyze the rate-limiting dehydration step and the need for *not too much* acid to avoid complete protonation of the amine. Each individual nucleophilic addition reaction has its own specific requirements, and reaction conditions must be controlled for maximum reaction rates to be obtained.

Imine formation from such reagents as hydroxylamine ( $\text{NH}_2\text{OH}$ ), semicarbazide ( $\text{NH}_2\text{NHCONH}_2$ ), and 2,4-dinitrophenylhydrazine are particularly useful because the products of these reactions—**oximes**, **semicarbazones**, and **2,4-dinitrophenylhydrazones** (2,4-DNPs)—are usually crystalline and easy to handle. Such crystalline derivatives are sometimes prepared as a means of identifying liquid ketones or aldehydes.

#### Oxime



#### Semicarbazone



#### 2,4-Dinitrophenylhydrazone



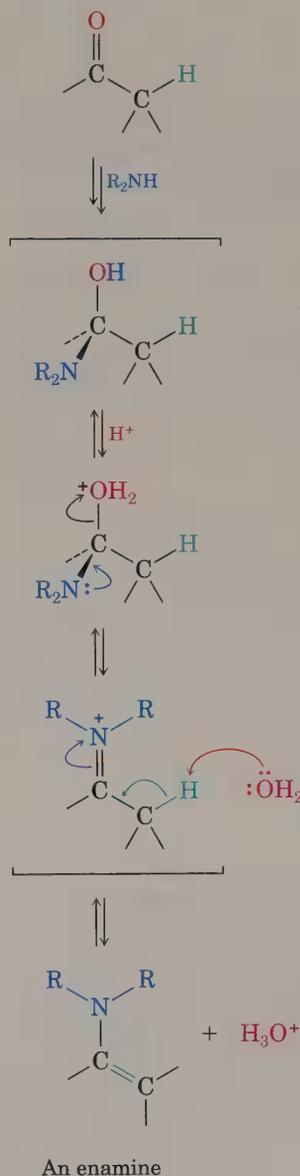
Enamines are formed when a ketone or an aldehyde reacts with a secondary amine,  $R_2NH$ . The process is identical to imine formation up to the iminium ion stage, but at this point there is no proton on nitrogen that can be lost to yield neutral product. Instead, a proton is lost from the alpha-carbon atom, yielding an enamine (Figure 19.10).

Nucleophilic addition of a secondary amine to the ketone or aldehyde, followed by proton transfer from nitrogen to oxygen, yields an intermediate carbinolamine in the normal way.

Protonation of the hydroxyl by acid catalyst converts it into a better leaving group.

Elimination of water by the lone-pair electrons on nitrogen then yields an intermediate iminium ion.

Loss of a proton from the alpha-carbon atom yields the enamine product and regenerates the acid catalyst.



**Figure 19.10** Mechanism of enamine formation by reaction of a ketone or an aldehyde with a secondary amine.

## PROBLEM.....

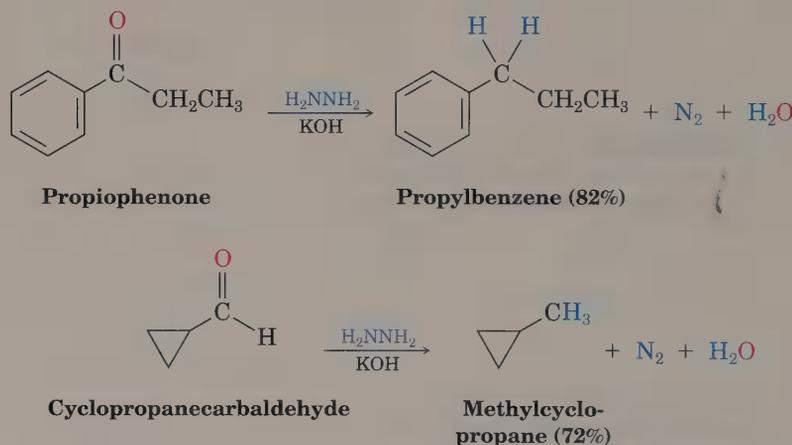
- 19.12 Show the products you would obtain by reaction of cyclohexanone with ethylamine,  $\text{CH}_3\text{CH}_2\text{NH}_2$ , and with diethylamine,  $(\text{CH}_3\text{CH}_2)_2\text{NH}$ .

## PROBLEM.....

- 19.13 Imine formation is reversible. Show all the steps involved in the hydrolysis of an imine to yield a ketone or an aldehyde plus primary amine.

## 19.13 Nucleophilic Addition of Hydrazine: The Wolff–Kishner Reaction

An important variant of the imine-forming reaction just discussed involves the treatment of a ketone or an aldehyde with hydrazine,  $\text{H}_2\text{NNH}_2$ , in the presence of  $\text{KOH}$ . This reaction, discovered independently in 1911 by Ludwig Wolff<sup>3</sup> in Germany and N. M. Kishner<sup>4</sup> in Russia, is an extremely valuable synthetic method for converting ketones or aldehydes into alkanes,  $\text{R}_2\text{C}=\text{O} \rightarrow \text{R}_2\text{CH}_2$ .



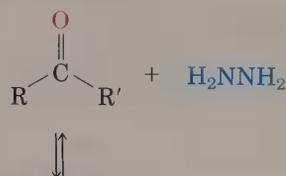
The **Wolff–Kishner reaction** is often carried out at  $240^\circ\text{C}$  in boiling diethylene glycol solvent, but a modification in which dimethyl sulfoxide is used as solvent allows the process to take place near room temperature.

The Wolff–Kishner reaction involves formation of a hydrazone intermediate, followed by base-catalyzed double-bond migration, loss of  $\text{N}_2$  gas, and formation of alkane product (Figure 19.11). The double-bond migration

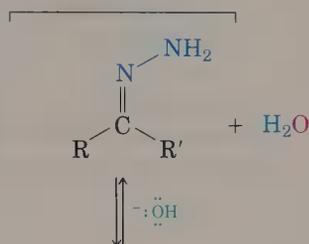
<sup>3</sup>Ludwig Wolff (1857–1919); b. Neustadt/Hardt; Ph.D. Strasbourg (Fittig); professor, University of Jena.

<sup>4</sup>N. M. Kishner (1867–1935); b. Moscow; Ph.D. Moscow (Markovnikov); professor, universities of Tomsk and Moscow.

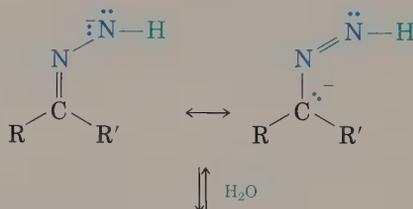
Reaction of the ketone or aldehyde with hydrazine yields a hydrazone in the normal way.



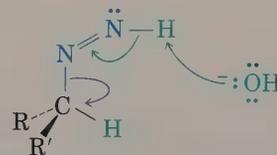
Base then abstracts one of the weakly acidic protons from  $-\text{NH}_2$ , yielding a hydrazone anion. This anion has an “allylic” resonance form that places the negative charge on carbon and the double bond between nitrogens.



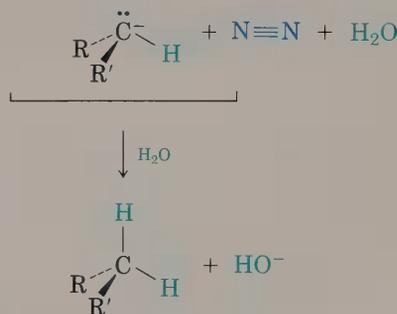
Protonation of the hydrazone anion takes place on carbon to yield a neutral intermediate.



Base-induced loss of nitrogen then gives a carbanion . . .



. . . that is protonated to yield neutral alkane product.



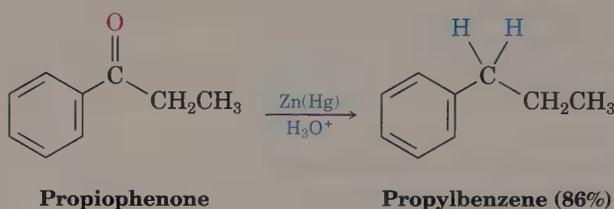
**Figure 19.11** Mechanism of the Wolff–Kishner reduction of a ketone or an aldehyde to yield an alkane.

takes place when base removes one of the weakly acidic NH protons to generate a hydrazone anion. Since the hydrazone anion has an allylic resonance structure that places the double bond between nitrogens and the negative charge on carbon, protonation can occur on carbon to generate the double-bond migrated product. The second-to-last step—loss of nitrogen to

give an alkyl anion that is then protonated—is an unusual reaction, driven primarily by the large thermodynamic stability of the  $N_2$  molecule.

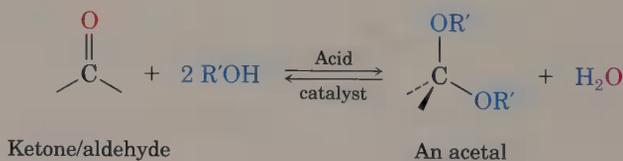
Note that the Wolff–Kishner reduction accomplishes the same overall transformation that we saw earlier during catalytic hydrogenation of acyl benzenes to alkylbenzenes (Section 16.12). The Wolff–Kishner reduction is more general and more useful than catalytic hydrogenation, however, since it works well with both alkyl and aryl ketones.

In addition to the Wolff–Kishner reaction, another process called the **Clemmensen<sup>5</sup> reduction** also accomplishes the conversion of a ketone or an aldehyde to an alkane. The Clemmensen reduction, whose mechanism is complex and not fully understood, involves treatment of the carbonyl compound with amalgamated zinc,  $Zn(Hg)$ , and concentrated aqueous  $HCl$ . It is used primarily when the ketone or aldehyde starting material is sensitive to the strongly basic conditions required by Wolff–Kishner reduction.



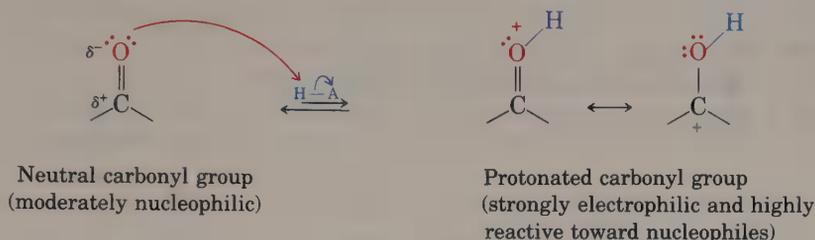
## 19.14 Nucleophilic Addition of Alcohols: Acetal Formation

Ketones and aldehydes react reversibly with alcohols in the presence of an acid catalyst to yield **acetals**,  $R_2C(OR')_2$ , also called *ketals* in older literature.

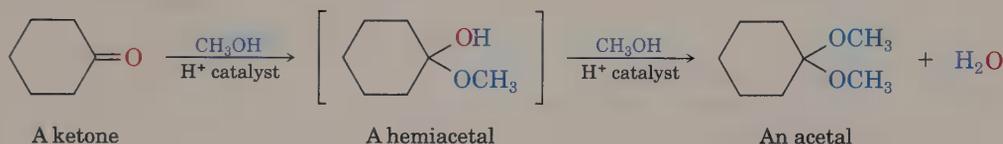


Acetal formation is similar to the hydration reaction studied in Section 19.8. Like water, alcohols are relatively weak nucleophiles that add to ketones and aldehydes only slowly under neutral conditions. Under acidic conditions, however, the nucleophilic carbonyl oxygen is protonated, and the resultant protonated carbonyl compound is much more reactive than its neutral counterpart. Addition of alcohol therefore occurs rapidly.

<sup>5</sup>E. C. Clemmensen (1876–1941); b. Odense, Denmark; Ph.D. Copenhagen; Clemmensen Chemical Corp., Newark, NY.

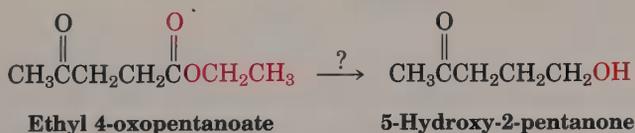


The initial nucleophilic addition of alcohol to the carbonyl group yields a hydroxy ether called a **hemiacetal**, analogous to the gem diol formed by addition of water (Section 19.8). Hemiacetals are formed reversibly, with the equilibrium normally favoring the carbonyl compound. In the presence of acid, however, a further reaction can occur. Protonation of the hydroxyl group, followed by an E1-like loss of water, leads to a cation (an **oxonium ion**,  $R_3O^+$ ) that then adds a second equivalent of alcohol to yield the acetal. For example, reaction of cyclohexanone with methanol yields the dimethyl acetal. The full mechanism is shown in Figure 19.12 on page 732.



Since all the steps of acetal formation are reversible, the reaction can be driven either forward (from carbonyl compound to acetal) or backward (from acetal to carbonyl compound), depending on the conditions chosen. The forward reaction is favored by choosing conditions that remove water from the medium and thus drive the equilibrium to the right. In practice, this is often done by distilling off water as it forms. The reverse reaction is favored by treating the acetal with mineral acid and a large excess of water to drive the equilibrium to the left.

Acetals are extremely useful compounds because they can serve as protecting groups for ketones and aldehydes in the same way that trimethylsilyl ethers serve as protecting groups for alcohols (Section 17.10). As we saw previously, it sometimes happens that one functional group might interfere with intended chemistry elsewhere in a complex molecule. For example, if we wanted to reduce only the ester group of ethyl 4-oxopentanoate, the ketone would interfere. Treatment of the starting keto ester with  $LiAlH_4$  would reduce both the keto and the ester groups to give a diol product:



By protecting the keto group as an acetal, however, the problem can be circumvented. Like other ethers, acetals are inert to bases, hydride reducing

Protonation of the carbonyl oxygen strongly polarizes the carbonyl group and . . .

. . . activates the carbonyl group for nucleophilic attack by oxygen lone-pair electrons from alcohol.

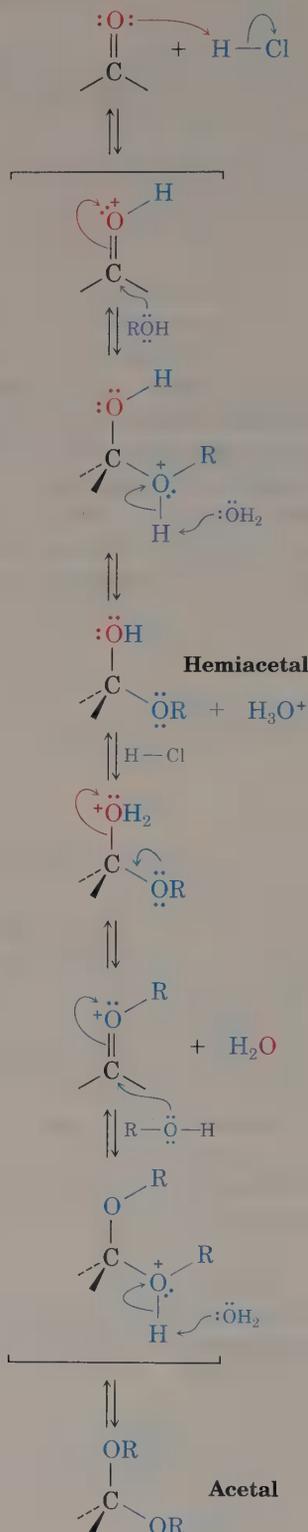
Loss of a proton yields a neutral hemiacetal tetrahedral intermediate.

Protonation of the hemiacetal hydroxyl converts it into a good leaving group.

Dehydration yields an intermediate oxonium ion.

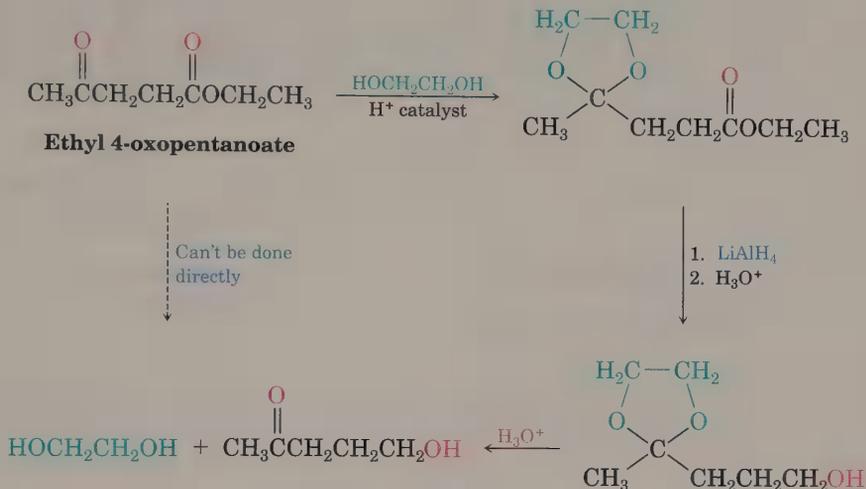
Addition of a second equivalent of alcohol gives protonated acetal.

Loss of a proton yields neutral acetal product.

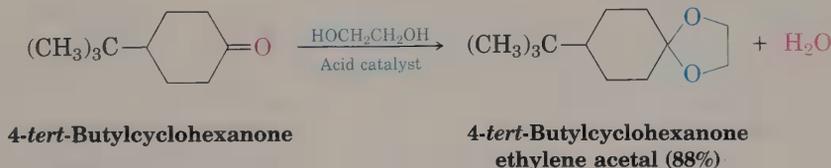


**Figure 19.12** Mechanism of acid-catalyzed acetal formation by reaction of a ketone or an aldehyde with an alcohol.

agents, Grignard reagents, and catalytic reducing conditions, but they're acid-sensitive. Thus, we can accomplish the selective reduction of the ester group in ethyl 4-oxopentanoate by first converting the keto group to an acetal, then reducing the ester with  $\text{LiAlH}_4$ , and then removing the acetal by treatment with aqueous acid.



In practice, it's convenient to use ethylene glycol as the alcohol and to form a cyclic acetal. The mechanism of cyclic acetal formation using one equivalent of ethylene glycol is exactly the same as that using two equivalents of methanol or other monoalcohol. The only difference is that both alcohol groups are now in the *same* molecule rather than in different molecules.



PROBLEM .....

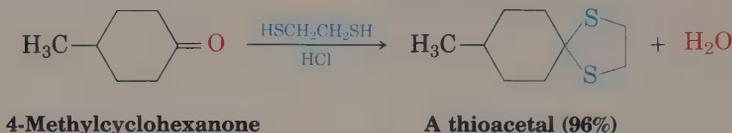
- 19.14** Show all of the steps in the acid-catalyzed formation of a cyclic acetal from ethylene glycol and a ketone or an aldehyde.

PROBLEM .....

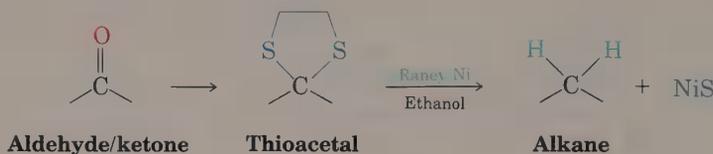
- 19.15** We saw in Section 19.7 that aldehydes are more reactive than ketones toward nucleophilic addition. How might you use this knowledge to carry out the following selective transformations? One of the two schemes requires a protection step.
- (a)  $\text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_2\text{CHO} \longrightarrow \text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
- (b)  $\text{CH}_3\text{COCH}_2\text{CH}_2\text{CH}_2\text{CHO} \longrightarrow \text{CH}_3\text{CH}(\text{OH})\text{CH}_2\text{CH}_2\text{CH}_2\text{CHO}$
- .....

## 19.15 Nucleophilic Addition of Thiols: Thioacetal Formation

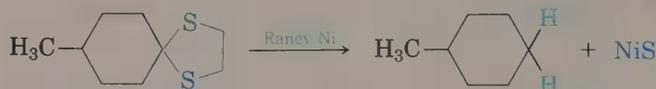
Thiols, RSH, add to ketones and aldehydes by a reversible, acid-catalyzed pathway to yield **thioacetals**,  $R'_2C(SR)_2$ . As you might expect, the mechanism of thioacetal formation is identical in all respects to that of acetal formation (Figure 19.12) except that a thiol is used in place of an alcohol. Ethanedithiol is often chosen, and the resultant cyclic thioacetals form rapidly in high yield.



Thioacetals are useful because they undergo *desulfurization* when treated with a specially prepared nickel powder known as **Raney<sup>6</sup> nickel** (Raney Ni). Since desulfurization removes sulfur from a molecule, replacing it by hydrogen, thioacetal formation followed by Raney nickel desulfurization is an excellent method for reducing ketones or aldehydes to alkanes.



For example,



Raney nickel desulfurization is a general method for reducing *any* R–S group to an R–H group. The mechanism of the process is not fully understood but undoubtedly involves radical intermediates. The hydrogen atoms in the desulfurized products come from hydrogen gas, which is adsorbed onto the Raney nickel surface during preparation.

Overall, the thioacetal desulfurization sequence is competitive with the Wolff–Kishner and Clemmensen reactions (Section 19.13) for accomplishing the reduction of a ketone or an aldehyde to an alkane. Note, however, that each of the three methods has its own advantages. The Wolff–Kishner reaction occurs under basic conditions, the Clemmensen reaction occurs under

<sup>6</sup>Murray Raney (1885–1966); b. Tennessee; B.A. University of Kentucky; D.Sc. (Hon.) University of Kentucky; Gilman Paint and Varnish Company (1925–1950); Raney Catalyst Company (1950–1966).

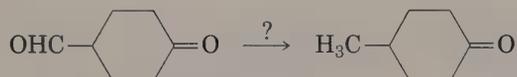
acidic conditions, and the Raney nickel desulfurization reaction occurs under neutral conditions.

PROBLEM .....

- 19.16 Show three different methods by which you might prepare cyclohexane from cyclohexanone.

PROBLEM .....

- 19.17 How might you use Raney nickel desulfurization of a thioacetal to carry out the following reduction?



## 19.16 Nucleophilic Addition of Phosphorus Ylides: The Wittig Reaction

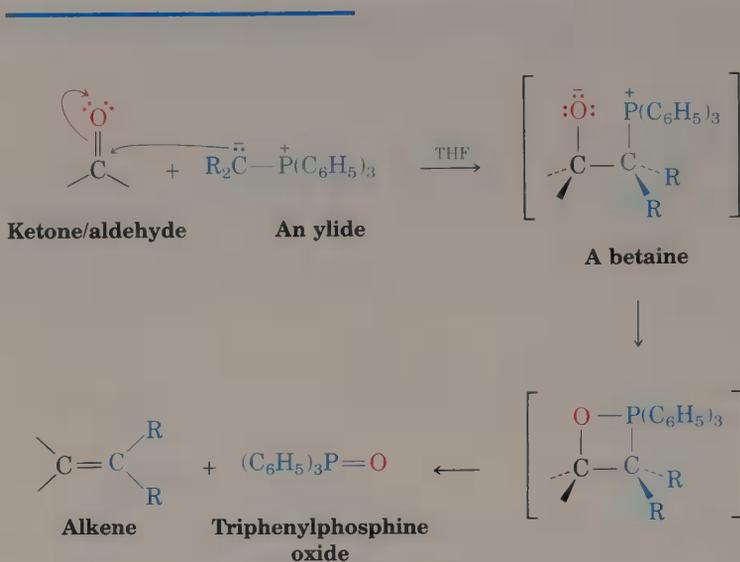
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Ketones and aldehydes are converted into alkenes by means of the **Wittig<sup>7</sup> reaction**. In this process, a phosphorus **ylide**,  $R_2C=P(C_6H_5)_3$  (also called a **phosphorane**), adds to a ketone or an aldehyde, yielding a dipolar intermediate called a **betaine**. (An ylide—pronounced **ill-id**—is a dipolar compound with adjacent plus and minus charges; a betaine—pronounced **bay-ta-teen**—is a dipolar compound in which the charges are nonadjacent.)

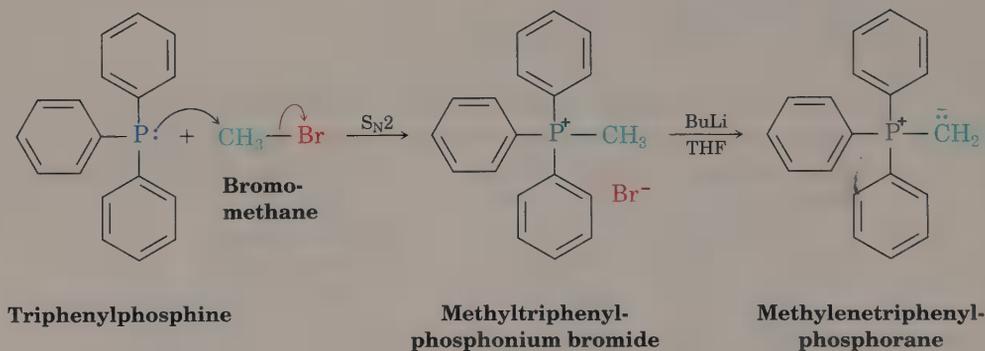
The betaine intermediate in the Wittig reaction is not isolated; rather, it decomposes to yield alkene and triphenylphosphine oxide. The net result is replacement of carbonyl oxygen by the organic fragment originally bonded to phosphorus (Figure 19.13).

The phosphorus ylides necessary for Wittig reaction are easily prepared by  $S_N2$  reaction of primary and some secondary (but not tertiary) alkyl halides with triphenylphosphine, followed by treatment with base. Triorganophosphines,  $R_3P$ , are generally excellent nucleophiles in  $S_N2$  reactions, and yields of crystalline tetraorganophosphonium salts are high. The proton on the carbon next to the positively charged phosphorus is weakly acidic and can be removed by a base such as sodium hydride or butyllithium (BuLi) to generate the neutral ylide. For example:

<sup>7</sup>Georg Wittig (1897–1987); b. Berlin; Ph.D. Marburg (von Auwers); professor, universities of Freiburg, Tübingen, and Heidelberg; Nobel Prize (1979).

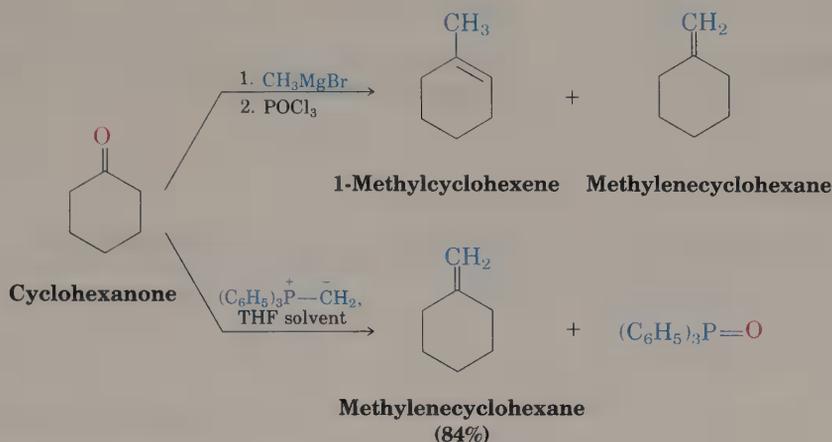


**Figure 19.13** The mechanism of the Wittig reaction between a ketone or an aldehyde and a phosphorus ylide.

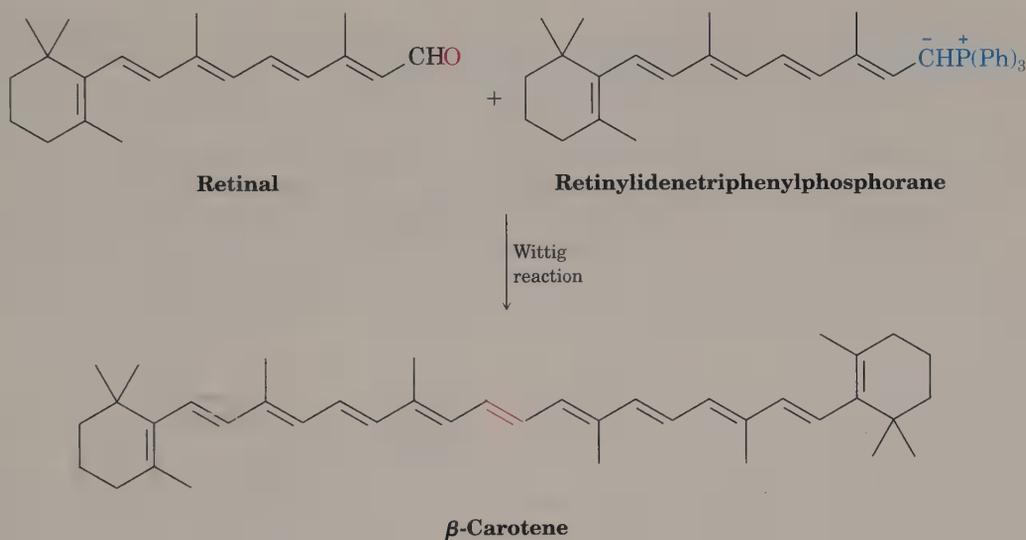


The Wittig reaction is extremely useful, and a great many mono-, di-, and trisubstituted alkenes can be prepared from the appropriate combination of phosphorane and ketone or aldehyde. Tetrasubstituted alkenes can't be prepared, however, presumably because of steric hindrance during the reaction.

The great value of the Wittig reaction is that pure alkenes of known structure are prepared. The alkene double bond is always exactly where the carbonyl group was in the precursor, and no product mixtures (other than *E,Z* isomers) are formed. For example, Wittig reaction of cyclohexanone with methylenetriphenylphosphorane yields only the single pure alkene product, methylenecyclohexane, whereas addition of methylmagnesium bromide to cyclohexanone followed by dehydration with POCl<sub>3</sub> yields a mixture of two alkenes, 1-methylcyclohexene and methylenecyclohexane.

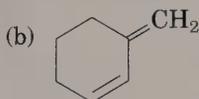
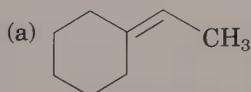


Wittig reactions are so important that they're even used commercially in a number of pharmaceutical applications. For example, the Swiss chemical firm of Hoffmann-LaRoche prepares  $\beta$ -carotene, a yellow food-coloring agent and dietary source of vitamin A, by Wittig reaction between retinal and retinylidenetriphenylphosphorane.



PROBLEM .....

19.18 What carbonyl compounds and what phosphorus ylides might you use to prepare these compounds?



- (c) 2-Methyl-2-hexene  
 (e) 1,2-Diphenylethylene

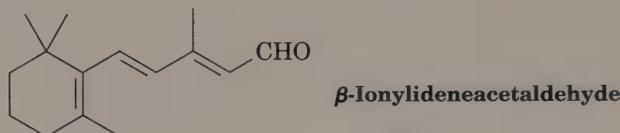


## PROBLEM .....

- 19.19 Why do you suppose triphenylphosphine, rather than, say, trimethylphosphine, is used to prepare Wittig reagents? What problems might you run into if trimethylphosphine were used?

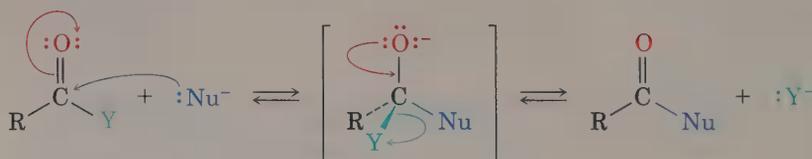
## PROBLEM .....

- 19.20 Another route to  $\beta$ -carotene involves a double Wittig reaction between two equivalents of  $\beta$ -ionylideneacetaldehyde and a diylide. Formulate the reaction, and show the structure of the diylide.



## 19.17 The Cannizzaro Reaction

We said in the overview of carbonyl chemistry that nucleophilic addition reactions are characteristic of ketones and aldehydes, but not of carboxylic acid derivatives. The reason for the difference is structural. As shown by the general reaction scheme in Figure 19.14, tetrahedral intermediates produced by addition of a nucleophile to a carboxylic acid derivative can eliminate a leaving group, leading to a net nucleophilic acyl substitution reaction. Tetrahedral intermediates produced by addition of a nucleophile to a ketone or an aldehyde have only alkyl and hydrogen substituents, however, and thus can't usually expel a leaving group. The **Cannizzaro**<sup>8</sup> reaction, discovered in 1853, is one exception to this rule.

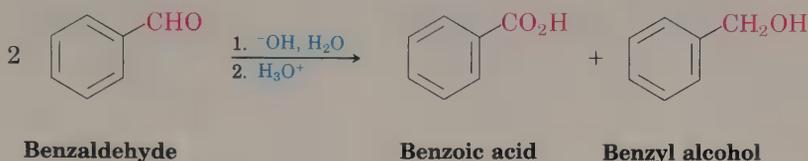


**Figure 19.14** The mechanism of nucleophilic acyl substitution reactions. Carboxylic acid derivatives have an electronegative group Y ( $-\text{Br}$ ,  $-\text{Cl}$ ,  $-\text{OR}'$ ,  $-\text{NH}_2$ ) that can act as a leaving group. Ketones and aldehydes have no such group.

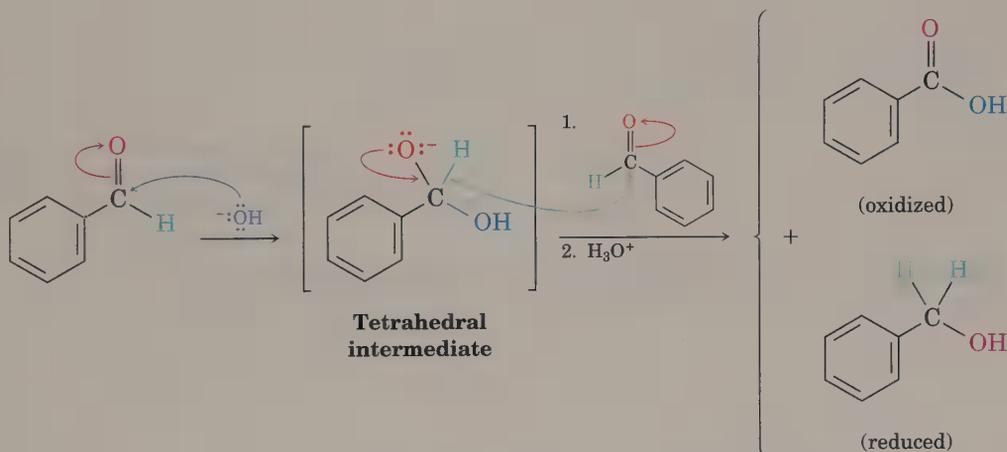
When an aldehyde such as benzaldehyde that has no hydrogens on the carbon next to the  $-\text{CHO}$  group is heated with hydroxide ion, a disproportion-

<sup>8</sup>Stanislao Cannizzaro (1826–1910); b. Palermo, Italy; studied at Pisa (Piria); professor, universities of Genoa, Palermo, and Rome.

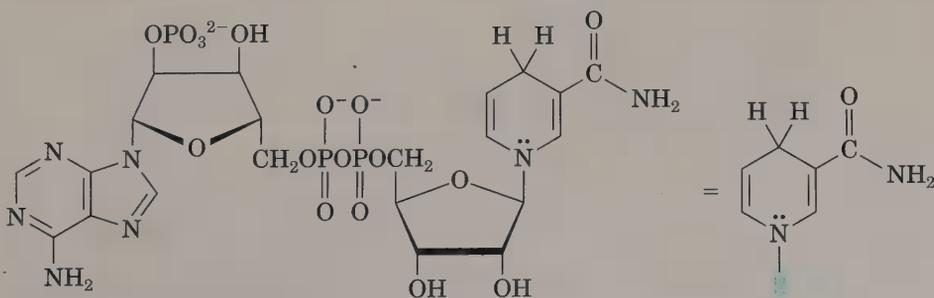
tionation reaction occurs, yielding one equivalent of carboxylic acid and one equivalent of alcohol.



The Cannizzaro reaction takes place by nucleophilic addition of hydroxide ion to the aldehyde to give a tetrahedral intermediate, *which expels hydride ion as a leaving group*. A second equivalent of aldehyde then accepts the hydride ion in another nucleophilic addition step. The net result is that one molecule of aldehyde undergoes an acyl substitution of hydroxide for hydride and is thereby oxidized to an acid, while a second molecule of aldehyde undergoes an addition of hydride and is thereby reduced to an alcohol.

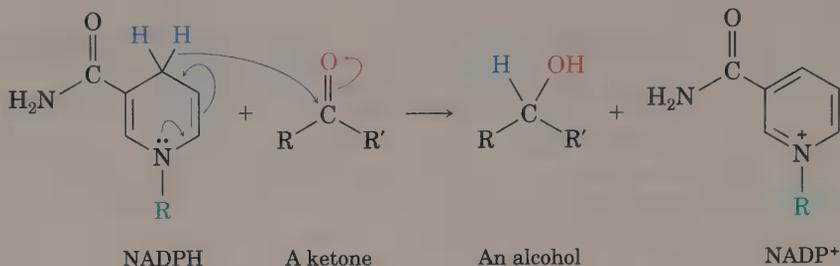


The Cannizzaro reaction has few practical applications and is effectively limited to formaldehyde and substituted benzaldehydes. It is interesting mechanistically, however, because it serves as a simple model for an important biological pathway by which reductions occur in living organisms. In nature, the most important reducing agent is a complex molecule named *nicotinamide adenine dinucleotide phosphate*, abbreviated NADPH.



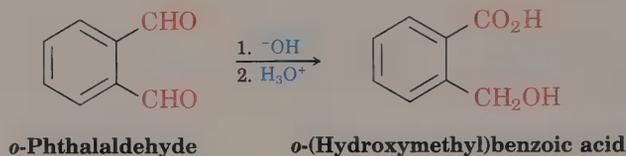
Nicotinamide adenine dinucleotide phosphate (NADPH)

NADPH functions as a hydride donor in much the same way the tetrahedral intermediate in a Cannizzaro reaction does. The electron lone pair on nitrogen expels  $\text{H}^-$ , which then adds to a carbonyl group in another molecule to effect a reduction. This is one of the key steps in the lipogenesis pathway by which organisms synthesize fatty acids.



PROBLEM .....

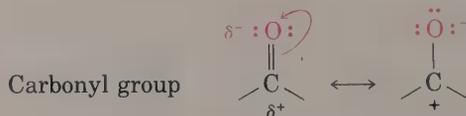
- 19.21 When *o*-phthalaldehyde is treated with base, *o*-(hydroxymethyl)benzoic acid is formed. Show the mechanism of this reaction.



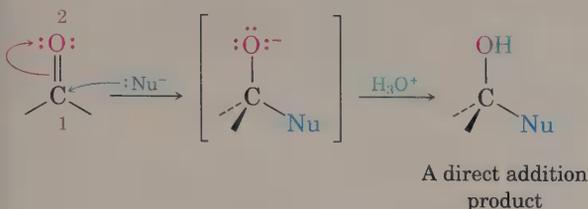
## 19.18 Conjugate Nucleophilic Addition to $\alpha,\beta$ -Unsaturated Carbonyl Groups

Addition of nucleophiles to carbonyl groups is one of the most important reactions in organic chemistry. Closely related to the direct additions we've been discussing is the **conjugate addition** of nucleophiles to  $\alpha,\beta$ -unsaturated ketones and aldehydes (Figure 19.15). The two processes are often distinguished by calling them 1,2 addition and 1,4 addition, respectively.

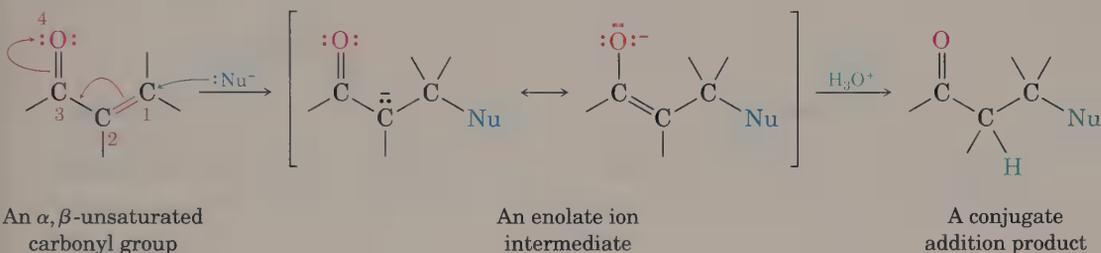
The conjugate addition of nucleophiles to  $\alpha,\beta$ -unsaturated carbonyl groups is caused by exactly the same electronic factors that are responsible for direct addition. We've seen that carbonyl groups are polarized so that the carbonyl carbon is electropositive, and we can even draw a dipolar resonance structure to underscore the point:



## Direct addition

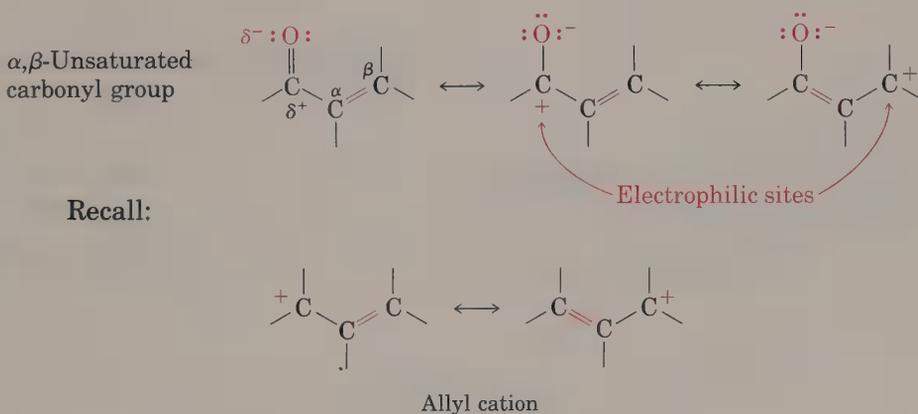


## Conjugate addition



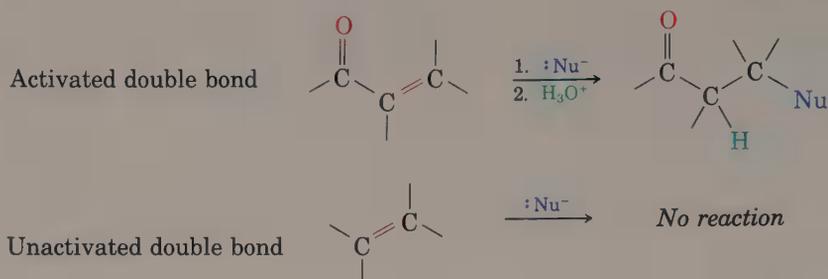
**Figure 19.15** A comparison of direct (1,2) and conjugate (1,4) nucleophilic addition reactions.

When a similar resonance structure is drawn for an  $\alpha,\beta$ -unsaturated carbonyl compound, however, the positive charge is part of an allylic cation and is shared by the beta carbon. In other words, the beta carbon of an  $\alpha,\beta$ -unsaturated ketone is an electrophilic site that can react with nucleophiles:



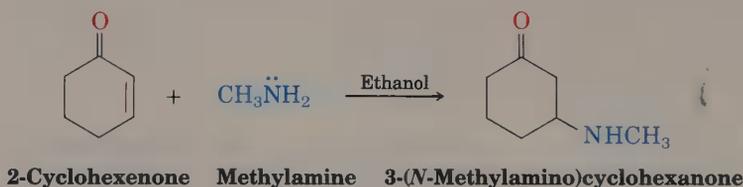
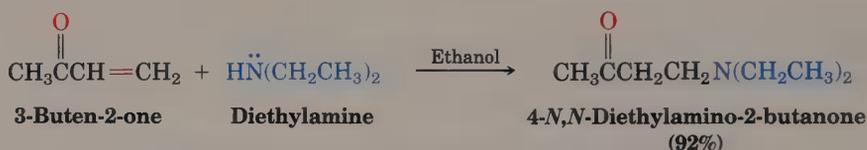
Conjugate addition of a nucleophile to the beta carbon of an enone leads to an enolate ion intermediate, which is then protonated on the alpha carbon to give the saturated ketone product (Figure 19.15). The net effect is addition of the nucleophile to the carbon-carbon double bond, with the carbonyl group itself unaffected. In fact, of course, the carbonyl group is crucial to

the success of the reaction. The carbon-carbon double bond would not be polarized, and no reaction would occur, without the carbonyl group.



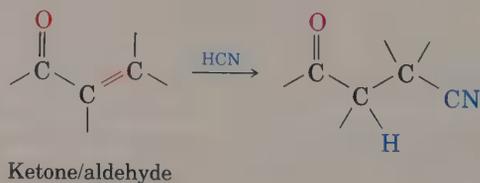
## CONJUGATE ADDITION OF AMINES

Primary and secondary amines add to  $\alpha,\beta$ -unsaturated carbonyl compounds, yielding  $\beta$ -amino ketones and aldehydes. Reaction occurs rapidly under mild conditions, and yields are generally excellent. Note that, if only one equivalent of amine is used, the conjugate addition product is obtained to the complete exclusion of the direct addition product.



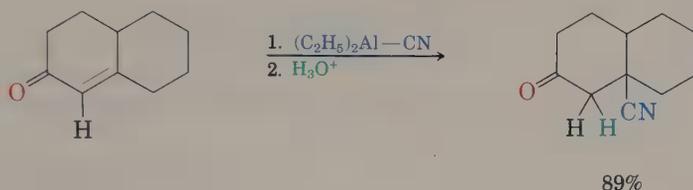
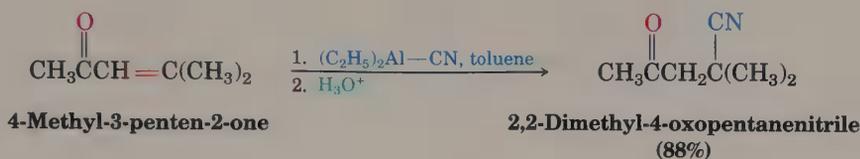
## CONJUGATE ADDITION OF HCN

The elements of HCN can be added to  $\alpha,\beta$ -unsaturated ketones and aldehydes, giving saturated keto nitriles:



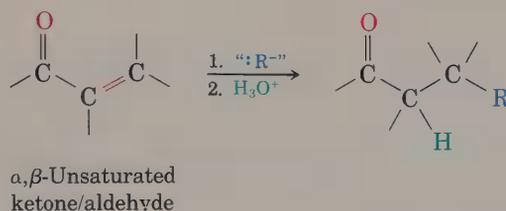
Although this process can be carried out using sodium cyanide in aqueous alcohol, higher yields are obtained using a method introduced in 1966

by Wataru Nagata.<sup>9</sup> The Nagata procedure for conjugate addition of HCN involves the use of diethylaluminum cyanide as the active reagent. Yields are generally excellent, as the following examples indicate. Note that only conjugate addition is observed; no cyanohydrin products of direct addition are formed.

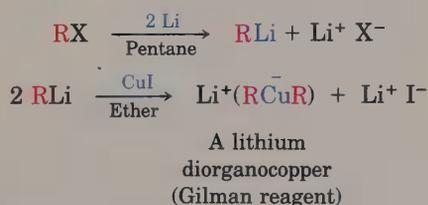


## CONJUGATE ADDITION OF ALKYL GROUPS: ORGANOCOPPER REACTIONS

Conjugate addition of an alkyl group to an  $\alpha,\beta$ -unsaturated ketone is one of the most important 1,4-addition reactions, just as direct addition of a Grignard reagent is one of the most important 1,2 additions.

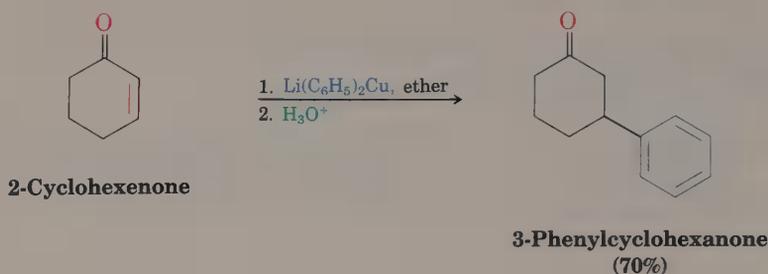
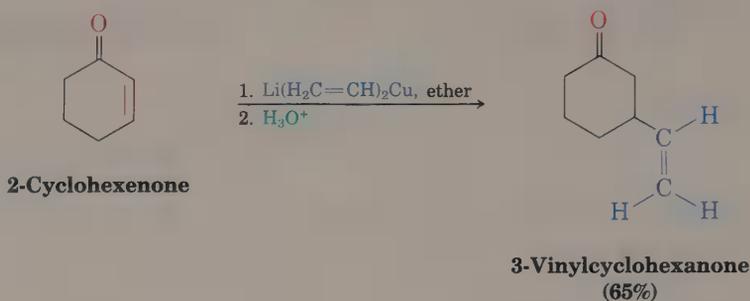
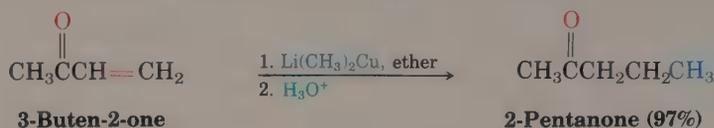


Conjugate addition of an alkyl group is carried out by treating the  $\alpha,\beta$ -unsaturated ketone with a lithium diorganocopper reagent. As we saw in Section 10.10, a wide variety of diorganocopper reagents can be prepared by reaction between one equivalent of cuprous iodide and two equivalents of organolithium:

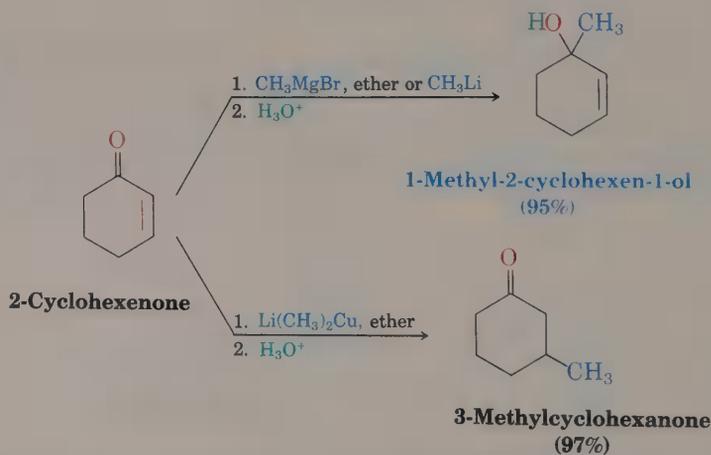


<sup>9</sup>Wataru Nagata (1922– ); b. Kohmoto, Hyogo Prefecture, Japan; Ph.D. University of Tokyo (Tsuda), 1961; Shionogi & Co. (1945–1988).

Primary, secondary, and even tertiary alkyl groups undergo the addition reaction, as do aryl and alkenyl groups. Alkynyl groups, however, react poorly in the conjugate addition process.



Diorganocopper reagents are unique in their ability to give conjugate addition products. Other organometallic reagents such as organomagnesiums (Grignard reagents) and organolithiums normally give direct carbonyl addition on reaction with  $\alpha,\beta$ -unsaturated ketones.

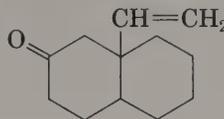


The mechanism of diorganocopper addition is not fully understood but probably involves radicals. The reaction is not a typical polar process like other nucleophilic additions are.

## PROBLEM

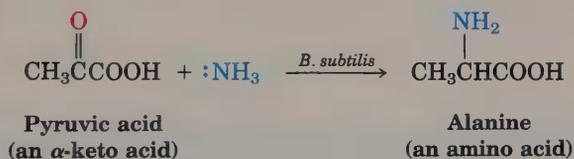
19.22 Show how conjugate addition reactions of lithium diorganocopper reagents might be used to synthesize these compounds:

- (a) 2-Heptanone  
 (b) 3,3-Dimethylcyclohexanone  
 (c) 4-*tert*-Butyl-3-ethylcyclohexanone  
 (d)

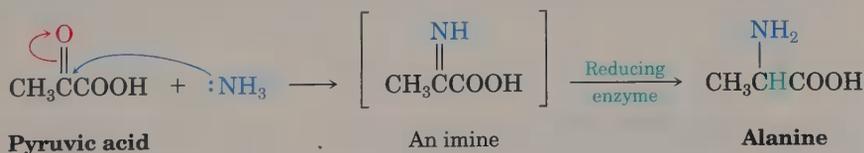


## 19.19 Some Biological Nucleophilic Addition Reactions

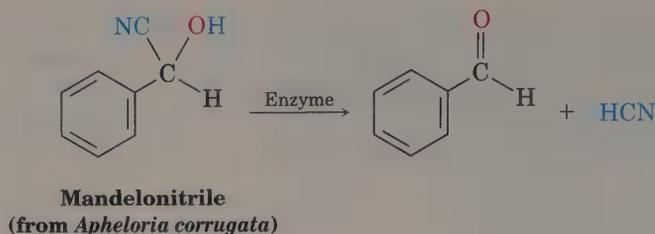
Nature synthesizes the molecules of life using many of the same reactions that chemists use in the laboratory. This is particularly true of carbonyl-group reactions, where nucleophilic addition steps play a critical role in the biosynthesis of many vital molecules. For example, one of the pathways by which amino acids are made in the body involves nucleophilic addition of ammonia to  $\alpha$ -keto acids. To choose one specific case, alanine is synthesized from pyruvic acid and ammonia by bacterial enzymes from *Bacillus subtilis*:



The key step in this biological transformation is the nucleophilic addition of ammonia to the ketone carbonyl group of pyruvic acid. The tetrahedral intermediate loses water to yield an imine, which is further reduced (a second nucleophilic addition step) by an enzymatic reaction to yield alanine.



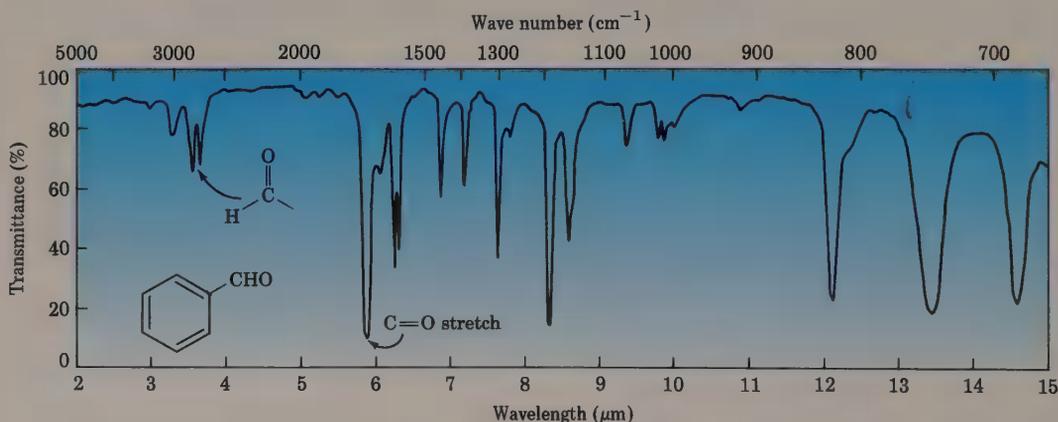
Another nucleophilic addition reaction—this time in reverse—plays an interesting role in the chemical defense mechanisms used by certain plants and insects to protect themselves against predators. When the millipede *Apheloria corrugata* is attacked by ants, it secretes the cyanohydrin mandelonitrile and an enzyme that catalyzes the decomposition of mandelonitrile into benzaldehyde and HCN. The millipede actually protects itself by discharging poisonous HCN at would-be attackers.



## 19.20 Spectroscopic Analysis of Ketones and Aldehydes

### INFRARED SPECTROSCOPY

Ketones and aldehydes show a strong C=O bond absorption in the infrared region  $1660\text{--}1770\text{ cm}^{-1}$ , as the spectra of benzaldehyde and cyclohexanone (Figures 19.16 and 19.17) demonstrate. In addition, aldehydes show two characteristic C–H absorptions in the range  $2720\text{--}2820\text{ cm}^{-1}$ , due to stretching of the aldehyde –CO–H bond. The exact position of the C=O bond absorption varies slightly from compound to compound but is highly diagnostic of the exact nature of the carbonyl group. Table 19.3 shows the correlation between the infrared absorption maximum and carbonyl-group structure.



**Figure 19.16** Infrared spectrum of benzaldehyde.

As the data in Table 19.3 indicate, saturated aldehydes usually show carbonyl absorptions near  $1730\text{ cm}^{-1}$  in the infrared spectrum, but conjugation of the aldehyde to an aromatic ring or a double bond lowers the absorption by  $25\text{ cm}^{-1}$  to near  $1705\text{ cm}^{-1}$ . Saturated aliphatic ketones and cyclohexanones both absorb near  $1715\text{ cm}^{-1}$ , and conjugation with a double

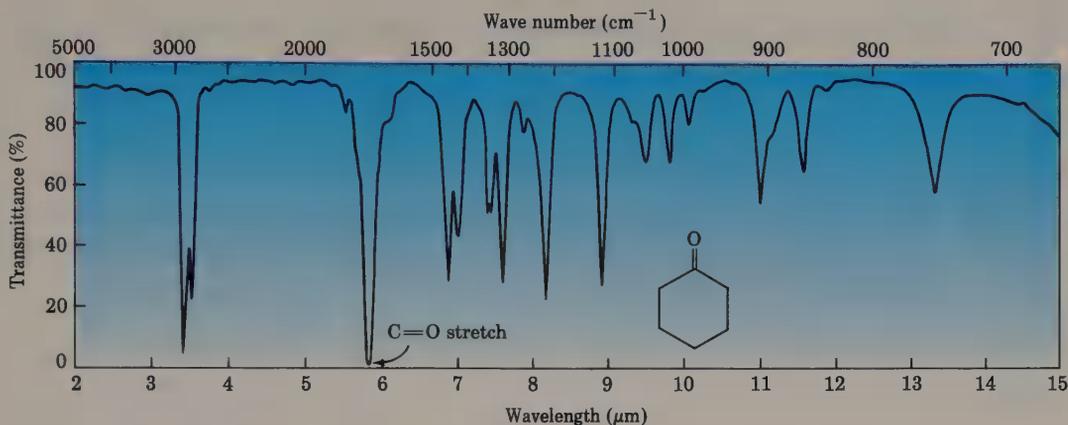
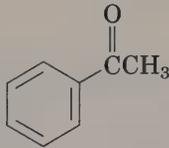


Figure 19.17 Infrared spectrum of cyclohexanone.

TABLE 19.3 Infrared Absorptions of Some Ketones and Aldehydes

Carbonyl type	Example	Infrared absorption ( $\text{cm}^{-1}$ )
Aliphatic aldehyde	Acetaldehyde	1730
Aromatic aldehyde	Benzaldehyde	1705
$\alpha,\beta$ -Unsaturated aldehyde	$\text{H}_2\text{C}=\text{CH}-\text{CHO}$	1705
Aliphatic ketone	Acetone	1715
Six-membered-ring ketone	Cyclohexanone	1715
Five-membered-ring ketone	Cyclopentanone	1750
Four-membered-ring ketone	Cyclobutanone	1785
Aromatic ketone		1690
$\alpha,\beta$ -Unsaturated ketone	$\text{H}_2\text{C}=\text{CHC}(=\text{O})\text{CH}_3$	1685

bond or an aromatic ring again lowers the absorption by  $30\text{ cm}^{-1}$  to  $1685\text{--}1690\text{ cm}^{-1}$ . Additional angle strain in the carbonyl group, caused by reducing the ring size of cyclic ketones to four or five, results in a marked raising of the absorption position.

The values given in Table 19.3 are remarkably constant from one ketone to another or from one aldehyde to another. As a result, infrared spectroscopy is an extraordinarily powerful tool for diagnosing the nature and chemical

environment of a carbonyl group in a molecule of unknown structure. An unknown that shows an infrared absorption at  $1730\text{ cm}^{-1}$  is almost certainly an aldehyde rather than a ketone; an unknown that shows an infrared absorption at  $1750\text{ cm}^{-1}$  is almost certainly a cyclopentanone, and so on.

PROBLEM .....

- 19.23** How might you use infrared spectroscopy to determine whether reaction between 2-cyclohexenone and lithium dimethylcopper gives the direct addition product or the conjugate addition product?

PROBLEM .....

- 19.24** Tell where you would expect each of these compounds to absorb in the infrared spectrum:
- |                                |                                  |
|--------------------------------|----------------------------------|
| (a) 4-Penten-2-one             | (b) 3-Penten-2-one               |
| (c) 2,2-Dimethylcyclopentanone | (d) <i>m</i> -Chlorobenzaldehyde |
| (e) 3-Cyclohexenone            | (f) 2-Hexenal                    |

PROBLEM .....

- 19.25** Dehydration of 3-hydroxy-3-phenylcyclohexanone by treatment with acid leads to an unsaturated ketone. What possible structures are there for the product? At what position in the infrared spectrum would you expect each to absorb? If the actual product has an absorption at  $1670\text{ cm}^{-1}$ , what is its structure?
- .....

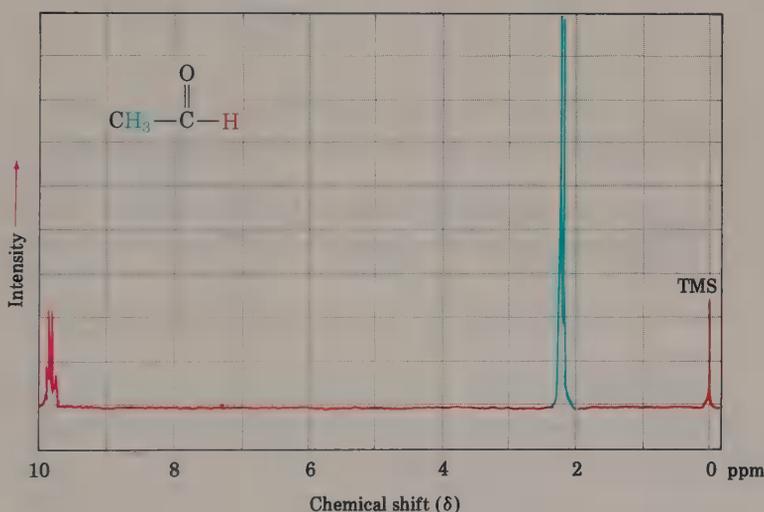
## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Carbonyl-group carbon atoms show readily identifiable and highly characteristic  $^{13}\text{C}$  NMR resonance peaks in the range  $190\text{--}215\ \delta$ . Since no other kinds of carbons absorb in this range, the presence of an NMR absorption near  $200\ \delta$  is strong evidence for a carbonyl group. Isolated ketone or aldehyde carbons usually absorb in the  $200\text{--}215\ \delta$  region, while  $\alpha,\beta$ -unsaturated carbonyl carbons absorb in the  $190\text{--}200\ \delta$  region. Table 19.4 lists some specific examples.

Proton NMR is also of considerable use for analysis of aldehydes, though less so for ketones. Aldehyde protons ( $\text{RCHO}$ ) absorb near  $10\ \delta$  in the  $^1\text{H}$  NMR spectrum and are highly distinctive since no other kind of proton absorbs in this region. The aldehyde proton usually shows spin-spin coupling to neighbor protons, with coupling constant  $J \approx 3\ \text{Hz}$ . Observation of the splitting pattern of the aldehyde proton tells the degree of substitution at the alpha position. Acetaldehyde, for example, shows a quartet at  $9.8\ \delta$  for the aldehyde proton, indicating that there are three protons neighboring the  $-\text{CHO}$  group (Figure 19.18).

**TABLE 19.4** Carbon-13 NMR Absorptions of Some Ketones and Aldehydes

Carbonyl compound	Carbon-13 NMR absorption of $\begin{array}{c} \diagup \\ \text{C}=\text{O} \ (\delta) \\ \diagdown \end{array}$
Acetaldehyde	201
Benzaldehyde	192
2-Butanone	207
Cyclohexanone	211
Acetophenone	196

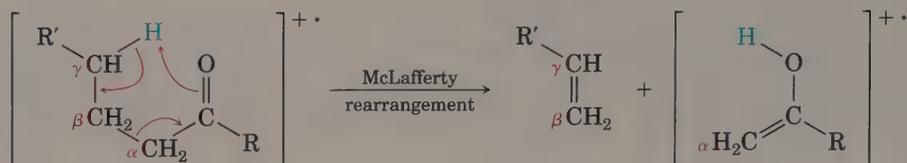
**Figure 19.18** Proton NMR spectrum of acetaldehyde.

Protons on the carbon next to a carbonyl group are slightly deshielded and normally absorb near 2.0–2.3  $\delta$ . (Note that the acetaldehyde methyl group in Figure 19.18 absorbs at 2.20  $\delta$ .) Methyl ketones are particularly distinctive, since they show a large sharp three-proton singlet near 2.1  $\delta$ . Complex spin–spin splittings often obscure the absorption patterns of other ketones, however, and reduce the diagnostic usefulness of  $^1\text{H}$  NMR.

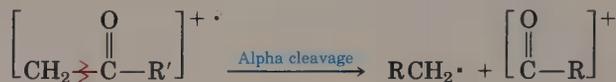
## MASS SPECTROSCOPY

Aliphatic ketones and aldehydes that have hydrogens on their gamma-carbon atoms undergo a characteristic mass spectral cleavage called the

**McLafferty<sup>10</sup> rearrangement.** A hydrogen atom is transferred from the gamma carbon to the carbonyl-group oxygen, the bond between the alpha and beta carbons is broken, and a neutral alkene fragment is produced. The charge remains with the oxygen-containing fragment.



In addition to fragmentation by the McLafferty rearrangement, ketones and aldehydes undergo cleavage of the bond between the alpha carbon and the carbonyl-group carbon, a so-called alpha-cleavage reaction. Alpha cleavage yields a neutral radical, while the charge remains with the oxygen-containing cation.



Fragment ions resulting from both alpha cleavage and McLafferty rearrangement can be seen in the mass spectrum of 5-methyl-2-hexanone in Figure 19.19. Alpha cleavage occurs primarily at the more substituted side of the carbonyl group, leading to a  $[\text{CH}_3\text{CO}]^+$  fragment with  $m/z = 43$ . McLafferty rearrangement and loss of 2-methylpropene yields a fragment with  $m/z = 58$ .

#### PROBLEM .....

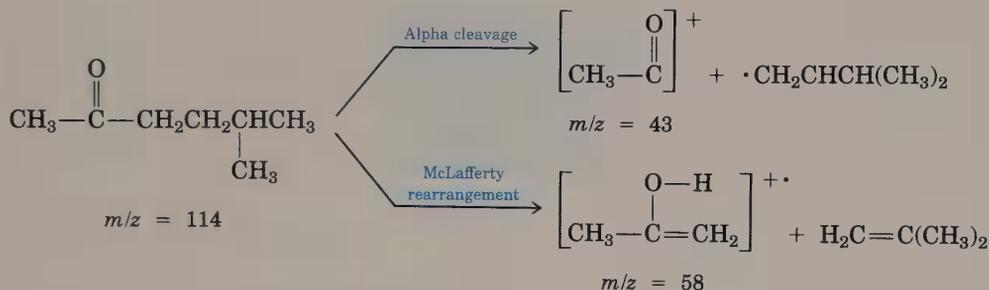
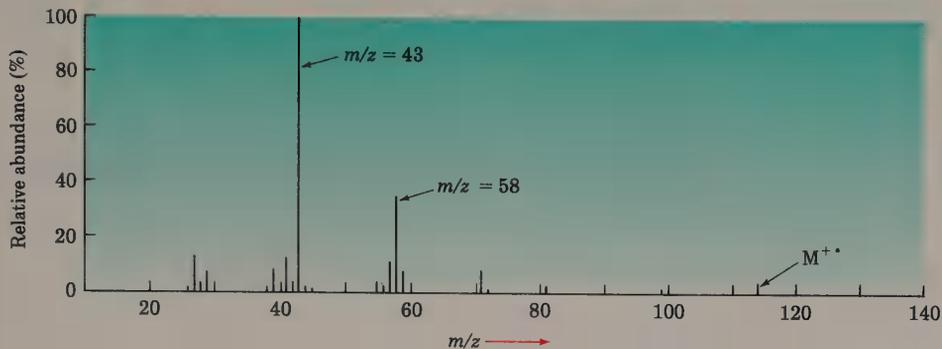
- 19.26 How might you use mass spectroscopy to distinguish between these pairs of isomers?
- 3-Methyl-2-hexanone and 4-methyl-2-hexanone
  - 3-Heptanone and 4-heptanone
  - 2-Methylpentanal and 4-methylpentanal
- .....

1-Methyl-2-cyclohexen-1-ol  
(95%)

## 19.21 Summary and Key Words

Aldehydes and ketones are among the most important of all compounds, both in biochemistry and in the chemical industry. Aldehydes are normally prepared in the laboratory by oxidative cleavage of alkenes, by oxidation of

<sup>10</sup>Fred Warren McLafferty (1923– ); b. Evanston, Ill.; Ph.D. Cornell University (1950); Dow Chemical (1950–1964); professor, Purdue University (1964–1968), Cornell University (1968– ).

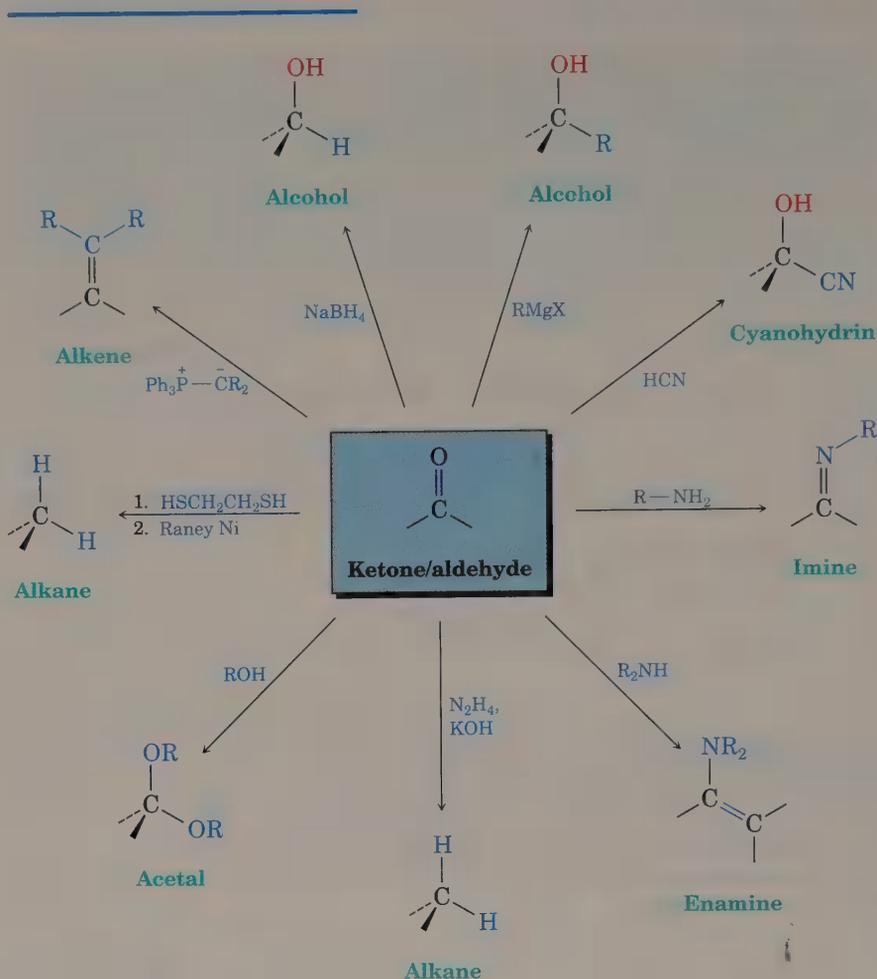


**Figure 19.19** Mass spectrum of 5-methyl-2-hexanone. The abundant peak at  $m/z = 43$  is due to alpha cleavage at the more highly substituted side of the carbonyl group. The peak at  $m/z = 58$  is due to McLafferty rearrangement. Note that the peak due to the molecular ion is very small.

primary alcohols with pyridinium chlorochromate (PCC), or by partial reduction of esters. Ketones are similarly prepared by oxidative cleavage of alkenes, by oxidation of secondary alcohols, or by addition of diorganocopper reagents to acid chlorides.

The **nucleophilic addition reaction** is the most important reaction of aldehydes and ketones. As shown in Figure 19.20 on the following page, many different product types can be prepared by nucleophilic additions. The reactions are applicable to ketones and aldehydes, but aldehydes are generally more reactive for both steric and electronic reasons.

Ketones and aldehydes are reduced by  $\text{NaBH}_4$  or  $\text{LiAlH}_4$  to yield secondary and primary alcohols, respectively. Addition of Grignard reagents to ketones and aldehydes also gives alcohols (tertiary and secondary, respectively), and addition of  $\text{HCN}$  yields cyanohydrins. Primary amines add to carbonyl compounds yielding **imines**, and secondary amines yield **enamines**. Reaction of a ketone or an aldehyde with hydrazine and base yields an alkane (the **Wolff-Kishner reaction**). Alcohols and thiols add to carbonyl groups to yield **acetals** and **thioacetals**, respectively. Acetals are valuable as protecting groups, and thioacetals are useful because they can be desulfurized by Raney nickel treatment to produce alkanes. Phosphoranes



**Figure 19.20** Some nucleophilic addition reactions of ketones and aldehydes.

add to ketones and aldehydes, giving alkenes (the **Wittig reaction**) in which the new C=C in the product is exactly where the C=O bond in the starting material was.

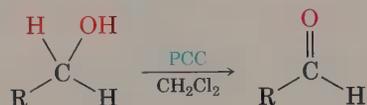
$\alpha,\beta$ -Unsaturated ketones and aldehydes often react with nucleophiles to give the product of **conjugate addition**, or **1,4 addition**. Particularly important is the reaction with diorganocopper reagents, which results in the addition of alkyl, aryl, and alkenyl (but not alkynyl) groups.

Ketones and aldehydes can be analyzed by infrared and nuclear magnetic resonance spectroscopy. Carbonyl groups absorb in the range 1660–1770  $\text{cm}^{-1}$ , with the exact infrared absorption position highly diagnostic of the precise kind of carbonyl group present in the molecule. Carbon-13 NMR spectroscopy is also useful for aldehydes and ketones since carbonyl carbons show resonances in the 190–215  $\delta$  range. Proton NMR is useful largely for analysis of aldehydes since aldehyde protons (R-CHO) absorb near 10  $\delta$ .

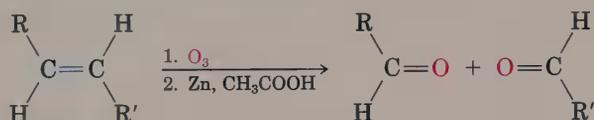
Ketones and aldehydes undergo two characteristic kinds of fragmentation in the mass spectrometer: **alpha cleavage** and **McLafferty rearrangement**.

## 19.22 Summary of Reactions

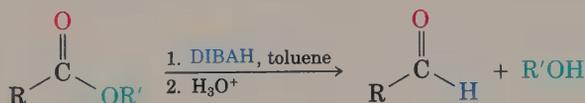
1. Preparation of aldehydes (Section 19.3)
  - a. Oxidation of primary alcohols (Section 17.9)



- b. Ozonolysis of alkenes (Section 7.9)



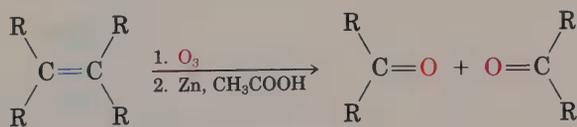
- c. Partial reduction of esters (Section 19.3)



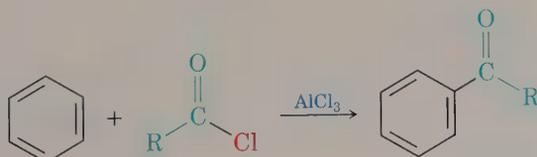
2. Preparation of ketones (Section 19.4)
  - a. Oxidation of secondary alcohols (Section 17.9)



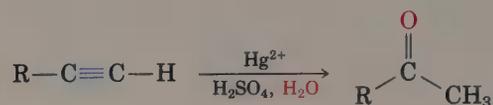
- b. Ozonolysis of alkenes (Section 7.9)



- c. Friedel–Crafts acylation (Section 16.4)



d. Alkyne hydration (Section 8.5)

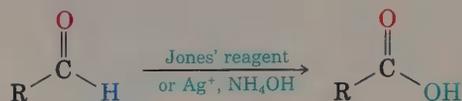


e. Diorganocopper reaction with acid chlorides (Section 19.4)

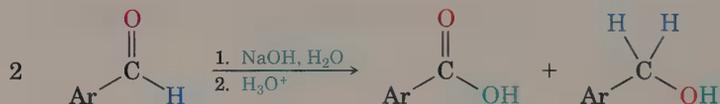


3. Reactions of aldehydes

a. Oxidation (Section 19.5)

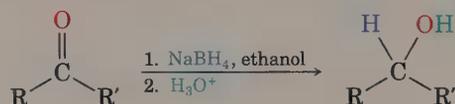


b. Cannizzaro reaction (Section 19.17)

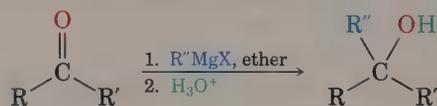


4. Nucleophilic addition reactions of aldehydes and ketones

a. Addition of hydride: reduction (Section 19.11)



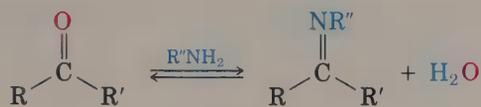
b. Addition of Grignard reagents (Section 19.10)



c. Addition of HCN: Cyanohydrins (Section 19.9)

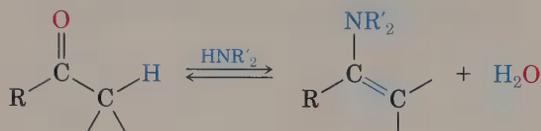


d. Addition of primary amines: Imines (Section 19.12)

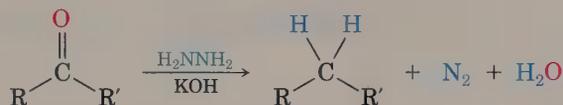


For example: oximes,  $\text{R}_2\text{C}=\text{N}-\text{OH}$   
 semicarbazones,  $\text{R}_2\text{C}=\text{N}-\text{NHCONH}_2$   
 2,4-dinitrophenylhydrazones,  
 $\text{R}_2\text{C}=\text{N}-\text{NH}-\text{C}_6\text{H}_4(\text{NO}_2)_2$

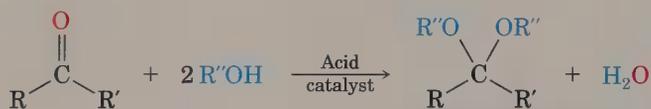
e. Addition of secondary amines: Enamines (Section 19.12)



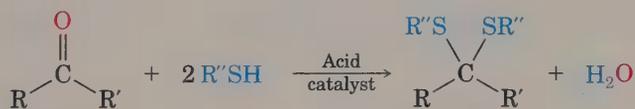
f. Wolff-Kishner reaction (hydrazine addition) (Section 19.13)



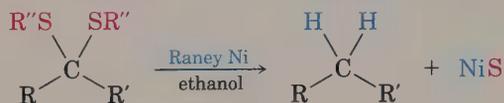
g. Addition of alcohols: Acetals (Section 19.14)



h. Addition of thiols: Thioacetals (Section 19.15)



i. Desulfurization of thioacetals with Raney nickel (Section 19.15)

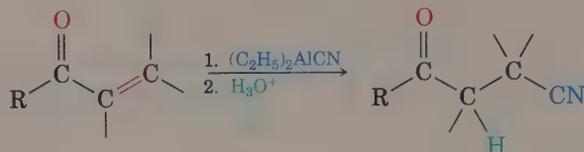


j. Addition of phosphorus ylides: Wittig reaction (Section 19.16)

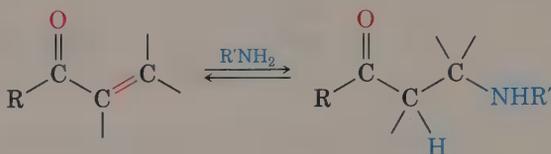


5. Conjugate additions to  $\alpha,\beta$ -unsaturated ketones and aldehydes (Section 19.18)

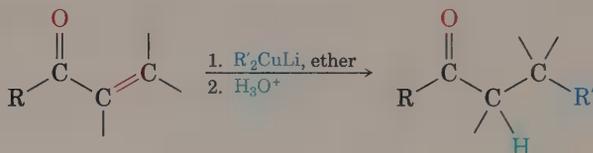
## a. Addition of HCN



## b. Addition of amines



## c. Addition of alkyl groups: Diorganocopper reaction



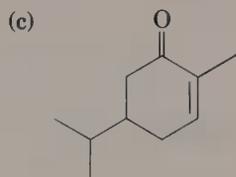
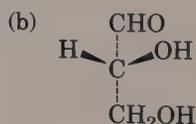
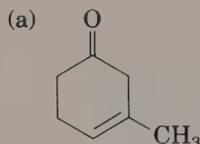
## ADDITIONAL PROBLEMS .....

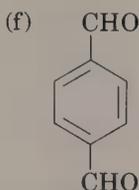
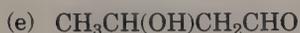
19.27 Draw structures corresponding to these names:

- |                                       |   |
|---------------------------------------|---|
| (a) Bromoacetone                      | (b) 3,5-Dinitrobenzaldehyde                           |
| (c) 2-Methyl-3-heptanone              | (d) 3,5-Dimethylcyclohexanone                         |
| (e) 2,2,4,4-Tetramethyl-3-pentanone   | (f) 4-Methyl-3-penten-2-one                           |
| (g) Butanedial                        | (h) 3-Phenyl-2-propenal                               |
| (i) 6,6-Dimethyl-2,4-cyclohexadienone | (j) <i>p</i> -Nitroacetophenone                       |
| (k) ( <i>S</i> )-2-Hydroxypropanal    | (l) (2 <i>S</i> ,3 <i>R</i> )-2,3,4-Trihydroxybutanal |

19.28 Draw and name the seven ketones and aldehydes having the formula  $\text{C}_5\text{H}_{10}\text{O}$ .

19.29 Provide IUPAC names for these structures:

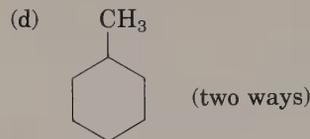
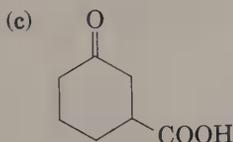
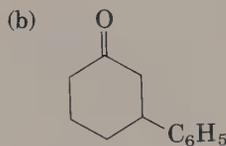
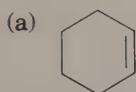
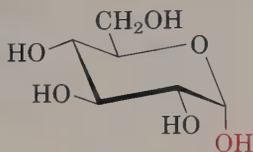


**19.30** Give structures that fit these descriptions:

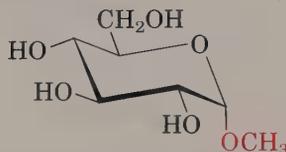
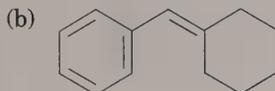
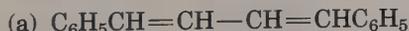
- (a) An  $\alpha,\beta$ -unsaturated ketone,  $\text{C}_6\text{H}_8\text{O}$       (b) An  $\alpha$ -diketone  
 (c) An aromatic ketone,  $\text{C}_9\text{H}_{10}\text{O}$       (d) A diene aldehyde,  $\text{C}_7\text{H}_8\text{O}$

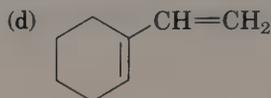
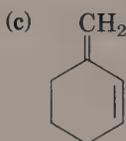
**19.31** Predict the products of the reaction of phenylacetaldehyde with these reagents:

- (a)  $\text{NaBH}_4$ , then  $\text{H}_3\text{O}^+$       (b) Tollens' reagent  
 (c) Hydroxylamine,  $\text{HCl}$       (d) Methylmagnesium bromide, then  $\text{H}_3\text{O}^+$   
 (e) Methanol plus acid catalyst      (f)  $\text{H}_2\text{NNH}_2/\text{KOH}$   
 (g) Methylene triphenylphosphorane      (h)  $\text{HCN}$ ,  $\text{KCN}$

**19.32** Answer Problem 19.31 for reaction with acetophenone.**19.33** How would you prepare the following substances from 2-cyclohexenone? More than one step may be required.**19.34** How can you account for the fact that glucose reacts with the Tollens reagent to give a silver mirror, but glucose  $\alpha$ -methyl glycoside does not?

Glucose

Glucose  $\alpha$ -methyl glycoside**19.35** Reaction of 2-butanone with  $\text{NaBH}_4$  yields a chiral product. What stereochemistry does the product have? Is it optically active?**19.36** Show how the Wittig reaction might be used to prepare these alkenes. Identify the alkyl halide and the carbonyl components that would be used.



19.37 How would you use a Grignard reaction on a ketone or an aldehyde to synthesize these compounds?

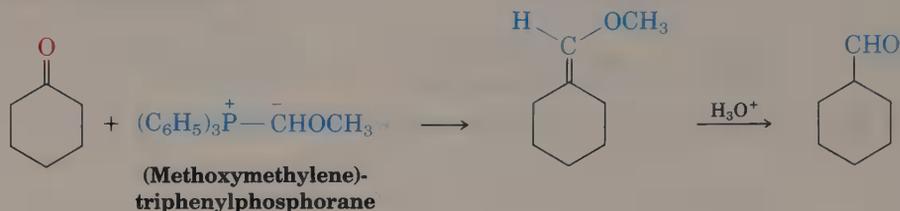
(a) 2-Pentanol

(b) 1-Butanol

(c) 1-Phenylcyclohexanol

(d) Diphenylmethanol

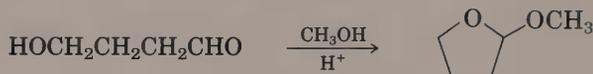
19.38 The Wittig reaction can be used to prepare aldehydes by using (methoxymethylene)triphenylphosphorane as the Wittig reagent and hydrolyzing the product with acid. For example:



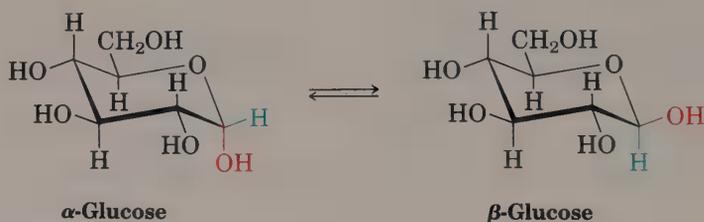
(a) How would you prepare the required phosphorane?

(b) Propose a mechanism for the hydrolysis step.

19.39 When 4-hydroxybutanal is treated with methanol in the presence of an acid catalyst, 2-methoxytetrahydrofuran is formed. Explain.

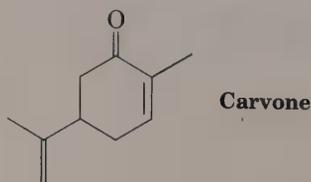


19.40 When crystals of pure  $\alpha$ -glucose are dissolved in water, isomerization slowly occurs to produce  $\beta$ -glucose. Propose a mechanism to explain this isomerization.



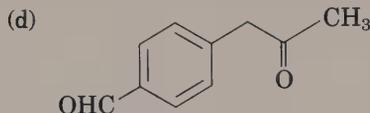
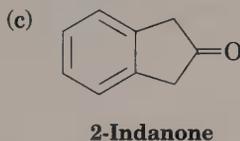
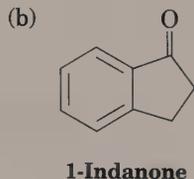
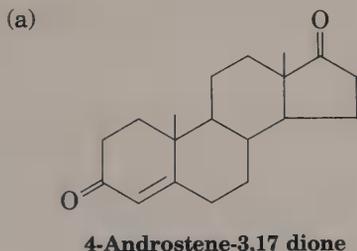
19.41 Give four methods for reducing a carbonyl group to a methylene group,  $R_2C=O \rightarrow R_2CH_2$ . What are the advantages and disadvantages of each?

19.42 Carvone is the major constituent of spearmint oil. What products would you expect from reaction of carvone with these reagents?

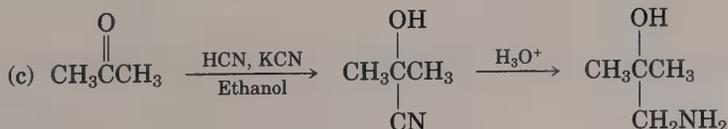
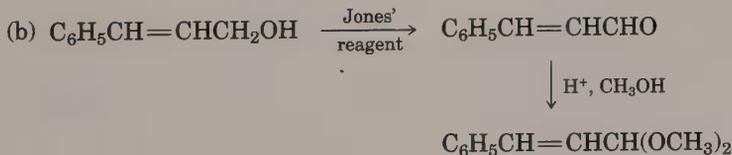
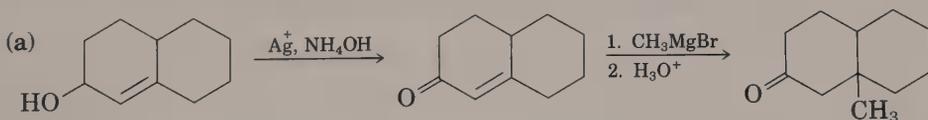


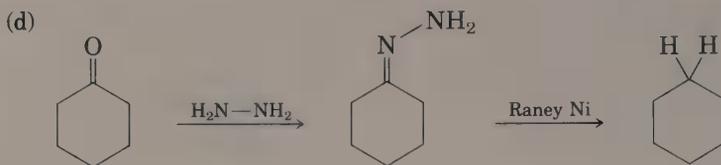
- (a)  $(\text{CH}_3)_2\text{Cu}^- \text{Li}^+$ , then  $\text{H}_3\text{O}^+$  (b)  $\text{LiAlH}_4$ , then  $\text{H}_3\text{O}^+$   
 (c)  $(\text{C}_2\text{H}_5)_2\text{AlCN}$ , then  $\text{H}_3\text{O}^+$  (d)  $\text{CH}_3\text{NH}_2$   
 (e)  $\text{C}_6\text{H}_5\text{MgBr}$ , then  $\text{H}_3\text{O}^+$  (f)  $\text{H}_2/\text{Pd}$   
 (g) Jones' reagent (h)  $(\text{C}_6\text{H}_5)_3\text{P}^+\text{CH}_2\text{CH}_3^-$   
 (i)  $\text{HSCH}_2\text{CH}_2\text{SH}$ , then Raney nickel (j)  $\text{HOCH}_2\text{CH}_2\text{OH}$ ,  $\text{HCl}$

- 19.43 Compound A, mol wt = 86, shows an infrared absorption at  $1730 \text{ cm}^{-1}$  and a very simple  $^1\text{H}$  NMR spectrum with peaks at  $9.7 \delta$  (1 H, singlet) and  $1.2 \delta$  (9 H, singlet). Propose a structure for A.
- 19.44 Compound B is isomeric with A (Problem 19.43) and shows an infrared peak at  $1715 \text{ cm}^{-1}$ . The  $^1\text{H}$  NMR spectrum of B has peaks at  $2.4 \delta$  (1 H, septet,  $J = 7 \text{ Hz}$ ),  $2.1 \delta$  (3 H, singlet), and  $1.2 \delta$  (6 H, doublet,  $J = 7 \text{ Hz}$ ). What is the structure of B?
- 19.45 How would you synthesize these compounds from cyclohexanone?  
 (a) 1-Methylcyclohexene (b) 2-Phenylcyclohexanone  
 (c) *cis*-1,2-Cyclohexanediol (d) 1-Cyclohexylcyclohexanol
- 19.46 At what position would you expect to observe infrared absorptions for the following molecules?



- 19.47 Each of the following reaction schemes contains one or more flaws. What is wrong in each case? How would you correct each scheme?



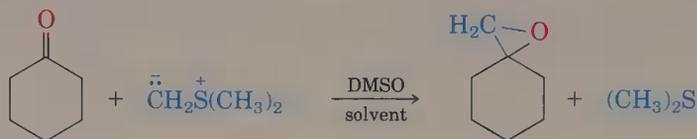


- 19.48 6-Methyl-5-hepten-2-one is a common constituent of many essential oils, particularly the lemongrass species. How could you synthesize this natural product from methyl 4-oxopentanoate?

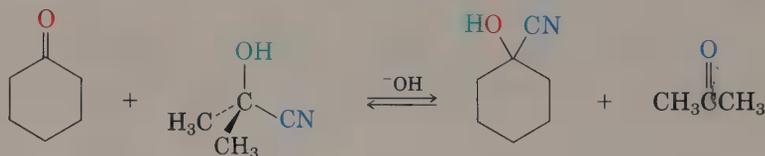


Methyl 4-oxopentanoate

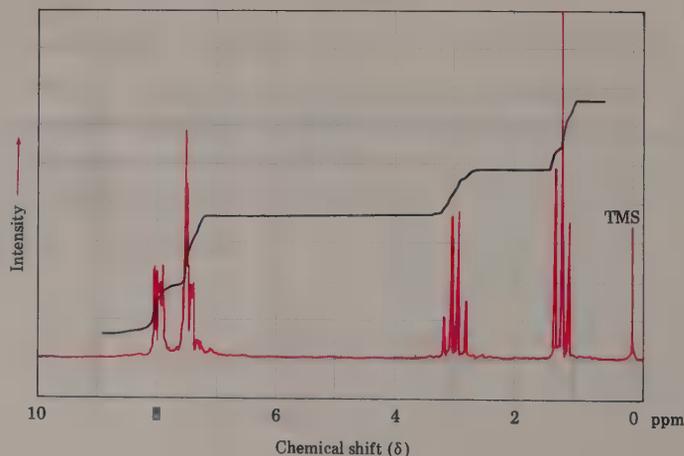
- 19.49 Ketones react with dimethylsulfonium methylide to yield epoxides. Suggest a mechanism for this reaction.

Dimethylsulfonium  
methylide

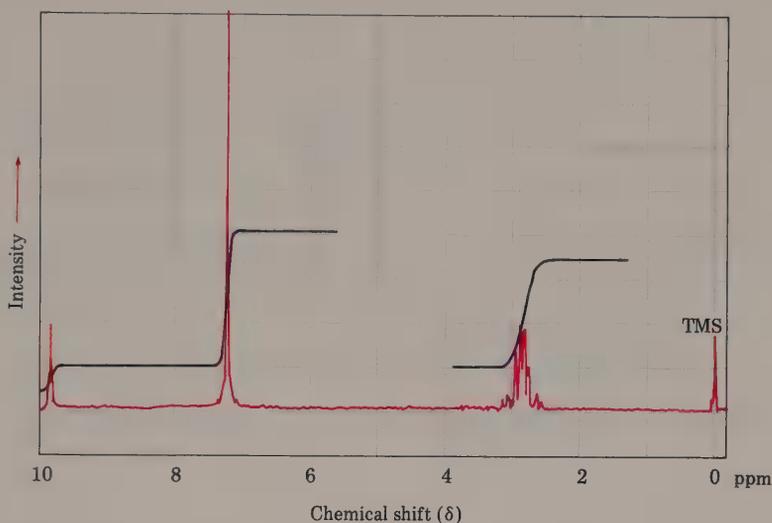
- 19.50 When cyclohexanone is heated in the presence of a large amount of acetone cyanohydrin and a small amount of base, cyclohexanone cyanohydrin and acetone are formed. Propose a mechanism for this transformation.



- 19.51 The NMR spectrum shown is that of a compound with formula  $\text{C}_9\text{H}_{10}\text{O}$ . How many double bonds and/or rings does this compound contain? If the unknown has an infrared absorption at  $1690\text{ cm}^{-1}$ , what is a likely structure?



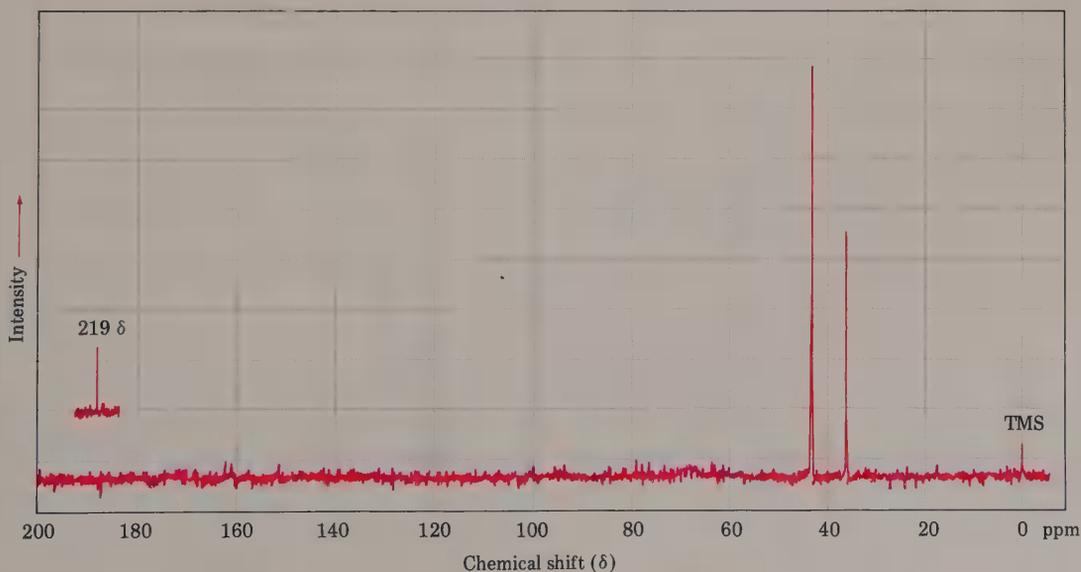
- 19.52** The NMR spectrum shown is that of a compound isomeric with the one in Problem 19.51. This isomer has an infrared absorption at  $1725\text{ cm}^{-1}$ . Propose a suitable structure.



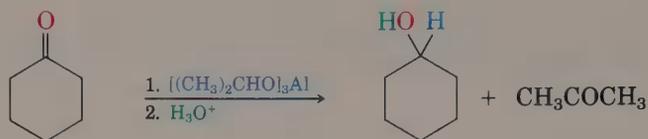
- 19.53** Propose structures for molecules that meet the following descriptions, where s = singlet, d = doublet, t = triplet, and q = quartet.

- (a)  $\text{C}_6\text{H}_{12}\text{O}$ ; IR:  $1715\text{ cm}^{-1}$ ;  
proton-coupled  $^{13}\text{C}$  NMR:  $8.0\ \delta$  (q),  $18.5\ \delta$  (q),  $33.5\ \delta$  (t),  $40.6\ \delta$  (d),  $214.0\ \delta$  (s)
- (b)  $\text{C}_5\text{H}_{10}\text{O}$ ; IR:  $1725\text{ cm}^{-1}$ ;  
proton-coupled  $^{13}\text{C}$  NMR:  $22.6\ \delta$  (q),  $23.6\ \delta$  (d),  $52.8\ \delta$  (t),  $202.4\ \delta$  (d)
- (c)  $\text{C}_6\text{H}_8\text{O}$ ; IR:  $1680\text{ cm}^{-1}$ ;  
proton-coupled  $^{13}\text{C}$  NMR:  $22.9\ \delta$  (t),  $25.8\ \delta$  (t),  $38.2\ \delta$  (t),  $129.8\ \delta$  (d),  $150.6\ \delta$  (d),  $198.7\ \delta$  (s)

- 19.54** Compound A,  $\text{C}_8\text{H}_{10}\text{O}_2$ , has an intense infrared absorption at  $1750\text{ cm}^{-1}$  and gives the  $^{13}\text{C}$  NMR spectrum shown. Propose a suitable structure for A.

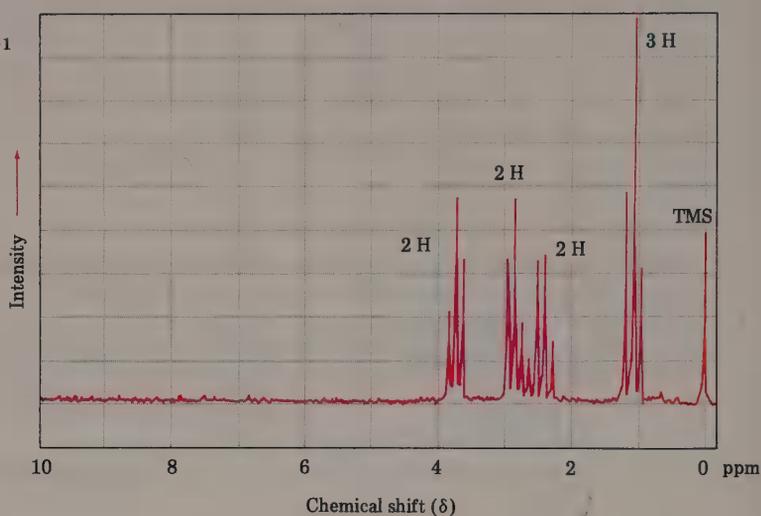


- 19.55 The Meerwein–Ponndorf–Verley reaction involves reduction of a ketone by treatment with an excess of aluminum triisopropoxide. The mechanism of the process is closely related to the Cannizzaro reaction, in which a hydride ion serves as leaving group. Formulate a reasonable mechanism.

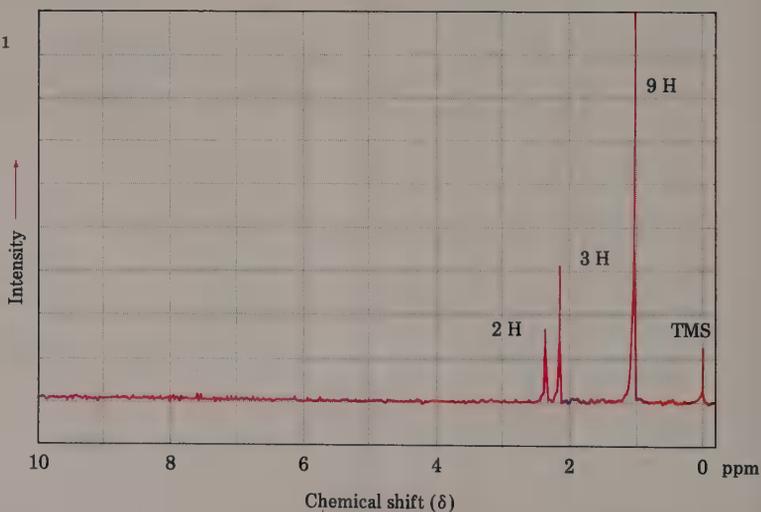


- 19.56 Propose structures for ketones or aldehydes that show the following  $^1\text{H}$  NMR spectra:

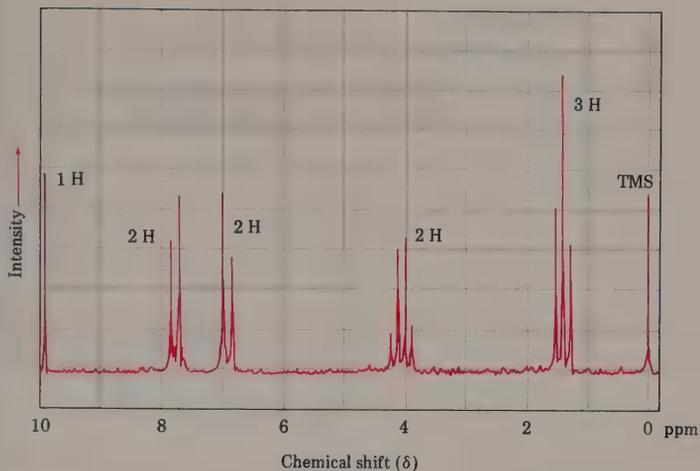
- (a)  $\text{C}_5\text{H}_9\text{ClO}$ ;  
IR:  $1710 \text{ cm}^{-1}$



- (b)  $\text{C}_7\text{H}_{14}\text{O}$ ;  
IR:  $1710 \text{ cm}^{-1}$

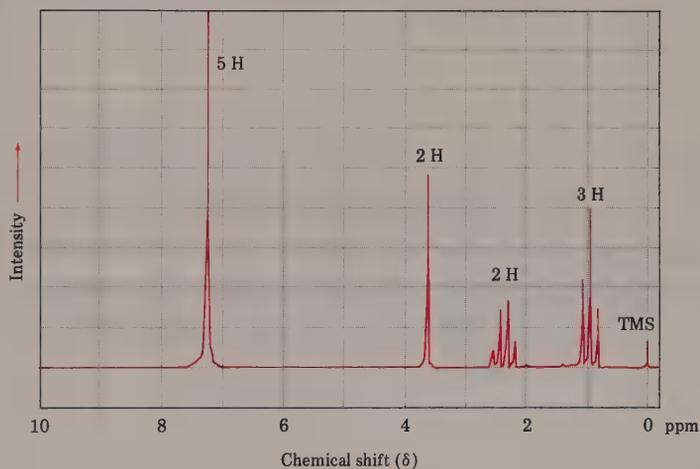


- (c)  $C_9H_{10}O_2$ ;  
IR:  $1695\text{ cm}^{-1}$

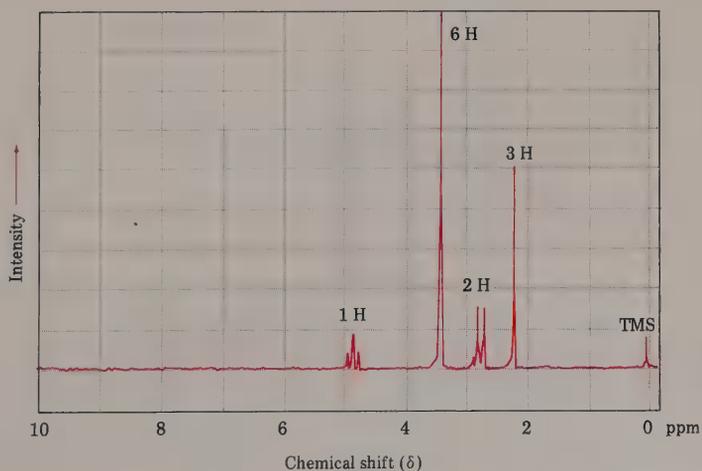


19.57 Propose structures for ketones or aldehydes that have the following  $^1\text{H}$  NMR spectra:

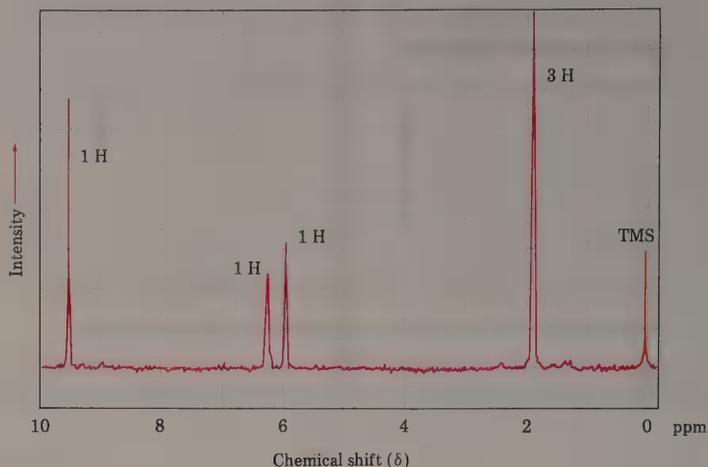
- (a)  $C_{10}H_{12}O$ ;  
IR:  $1710\text{ cm}^{-1}$



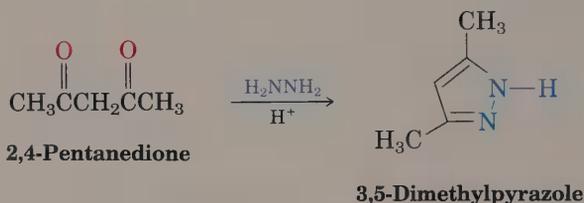
- (b)  $C_6H_{12}O_3$ ;  
IR:  $1715\text{ cm}^{-1}$



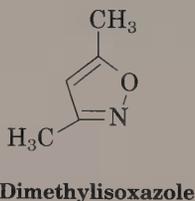
(c)  $C_4H_6O$ ;  
IR:  $1690\text{ cm}^{-1}$



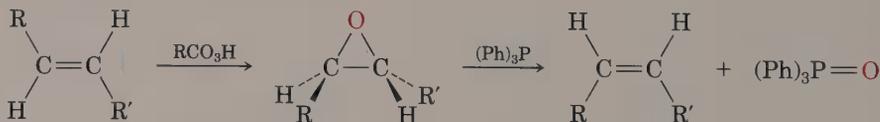
- 19.58 Propose a mechanism to account for the formation of 3,5-dimethylpyrazole from hydrazine and 2,4-pentanedione. Look carefully to see what has happened to each carbonyl carbon in going from starting material to product.

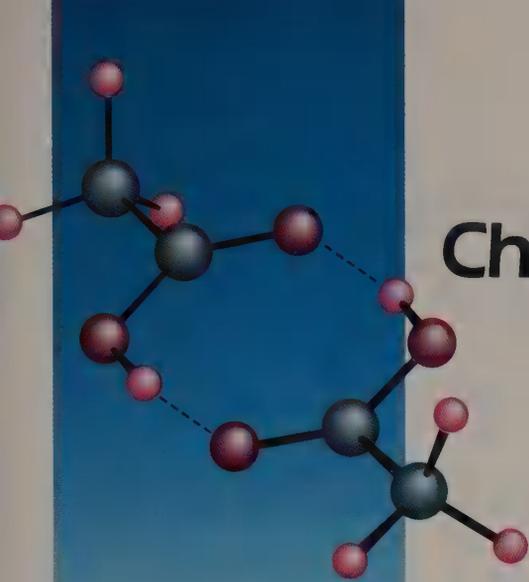


- 19.59 In light of your answer to Problem 19.58, propose a mechanism for the formation of 3,5-dimethylisoxazole from hydroxylamine and 2,4-pentanedione.



- 19.60 Trans alkenes are converted into their cis isomers and vice versa on epoxidation followed by treatment of the epoxide with triphenylphosphine. Propose a mechanism for the epoxide  $\rightarrow$  alkene reaction.

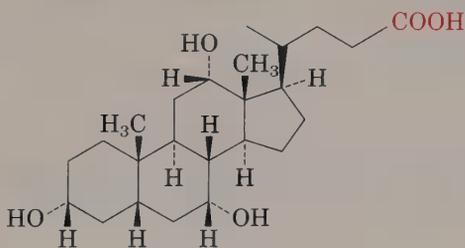




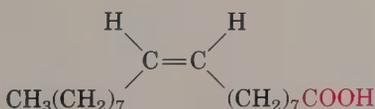
# Chapter 20

## Carboxylic Acids

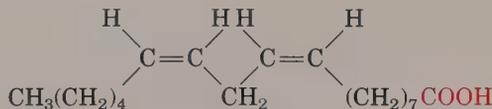
Carboxylic acids occupy a central place among acyl derivatives. Not only are they important compounds themselves, they also serve as building blocks for preparing related derivatives such as esters and amides. Among important examples are cholic acid, a major component of human bile, and long-chain aliphatic acids such as oleic acid and linoleic acid, which are biological precursors of fats and other lipids.



Cholic acid



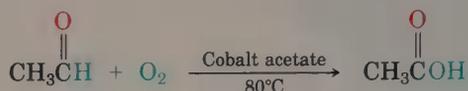
Oleic acid



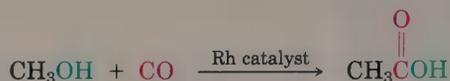
Linoleic acid

Many simple saturated carboxylic acids are also found in nature. For example, acetic acid,  $\text{CH}_3\text{COOH}$ , is the chief organic component of vinegar; butanoic acid,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ , is responsible for the rancid odor of sour butter; and hexanoic acid (caproic acid),  $\text{CH}_3(\text{CH}_2)_4\text{COOH}$ , is responsible for the unmistakable aroma of goats and dirty gym socks (Latin *capra*, "goat").

Some 1.6 million tons per year of acetic acid are produced industrially for a variety of purposes, including use as a raw material for preparing the vinyl acetate polymer used in paints and adhesives. The industrial method of acetic acid synthesis involves a cobalt acetate-catalyzed air oxidation of acetaldehyde, but this method is not used in the laboratory.

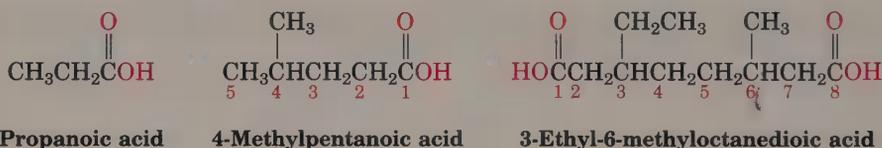


The Monsanto Company has developed an even more efficient synthesis involving a direct rhodium-catalyzed carbonylation of methanol:

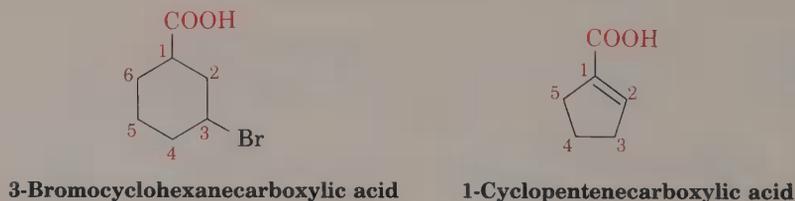


## 20.1 Nomenclature of Carboxylic Acids

IUPAC rules allow for two systems of nomenclature, depending on the complexity of the acid molecule. Carboxylic acids that are derived from open-chain alkanes are systematically named by replacing the terminal *-e* of the corresponding alkane name with *-oic acid*. The carboxyl carbon atom is always numbered C1 in this system.



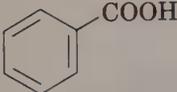
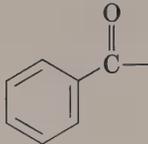
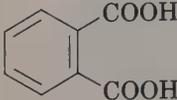
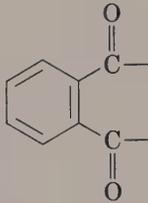
Alternatively, compounds that have a  $-\text{COOH}$  group bonded to a ring are named using the suffix *-carboxylic acid*. The carboxylic acid carbon is attached to C1 and is not itself numbered in this system.



For historical reasons having to do with the fact that many carboxylic acids were among the first organic compounds to be isolated and purified, IUPAC rules make allowance for a large number of well-entrenched common names, some of which are given in Table 20.1. We'll use systematic names in this book, with the exception of formic (methanoic) acid and acetic (ethanoic) acid, whose names are so well known that it makes little sense to refer

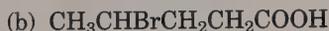
to them any other way. Also listed in Table 20.1 are the common names used for acyl groups derived from the parent acids.

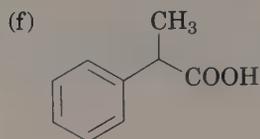
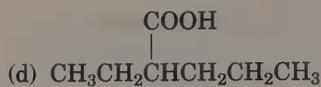
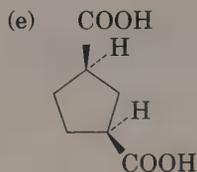
**TABLE 20.1** Some Common Names of Carboxylic Acids and Acyl Groups

<i>Carboxylic acid</i>		<i>Acyl group</i>	
<i>Structure</i>	<i>Name</i>	<i>Name</i>	<i>Structure</i>
HCOOH	Formic	Formyl	HCO—
CH <sub>3</sub> COOH	Acetic	Acetyl	CH <sub>3</sub> CO—
CH <sub>3</sub> CH <sub>2</sub> COOH	Propionic	Propionyl	CH <sub>3</sub> CH <sub>2</sub> CO—
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	Butyric	Butyryl	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>2</sub> CO—
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	Valeric	Valeryl	CH <sub>3</sub> (CH <sub>2</sub> ) <sub>3</sub> CO—
(CH <sub>3</sub> ) <sub>3</sub> CCOOH	Pivalic	Pivaloyl	(CH <sub>3</sub> ) <sub>3</sub> CCO—
HOOC <sub>2</sub> COOH	Oxalic	Oxalyl	—OCCO—
HOOCCH <sub>2</sub> COOH	Malonic	Malonyl	—OCCH <sub>2</sub> CO—
HOOCCH <sub>2</sub> CH <sub>2</sub> COOH	Succinic	Succinyl	—OC(CH <sub>2</sub> ) <sub>2</sub> CO—
HOOCCH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	Glutaric	Glutaryl	—OC(CH <sub>2</sub> ) <sub>3</sub> CO—
HOOCCH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	Adipic	Adipoyl	—OC(CH <sub>2</sub> ) <sub>4</sub> CO—
H <sub>2</sub> C=CHCOOH	Acrylic	Acryloyl	H <sub>2</sub> C=CHCO—
CH <sub>3</sub> CH=CHCOOH	Crotonic	Crotonoyl	CH <sub>3</sub> CH=CHCO—
H <sub>2</sub> C=C(CH <sub>3</sub> )COOH	Methacrylic	Methacryloyl	H <sub>2</sub> C=C(CH <sub>3</sub> )CO—
HC≡CCOOH	Propiolic	Propiolyyl	HC≡CCO—
HOOCCH=CHCOOH	<i>cis</i> -Maleic	Maleoyl	—OCCH=CHCO—
	<i>trans</i> -Fumaric	Fumaroyl	
	Benzoic	Benzoyl	
	Phthalic	Phthaloyl	

**PROBLEM** .....

**20.1** Give IUPAC names for these compounds:





## PROBLEM .....

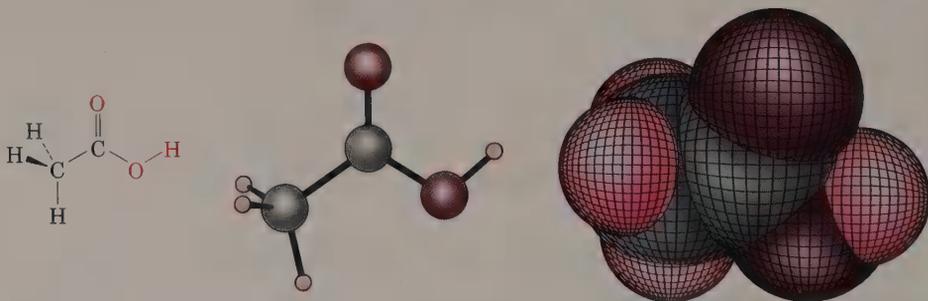
**20.2** Draw structures corresponding to these IUPAC names:

- |  |                                   |
|--|-----------------------------------|
| (a) 2,3-Dimethylhexanoic acid                            | (b) 4-Methylpentanoic acid        |
| (c) <i>trans</i> -1,2-Cyclobutanedicarboxylic acid       | (d) <i>o</i> -Hydroxybenzoic acid |
| (e) (9 <i>Z</i> ,12 <i>Z</i> )-9,12-Octadecadienoic acid |                                   |

## 20.2 Structure and Physical Properties of Carboxylic Acids

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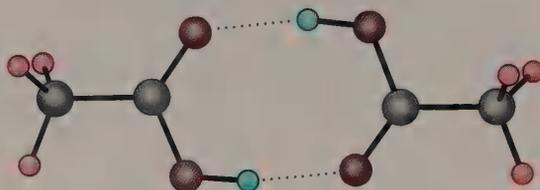
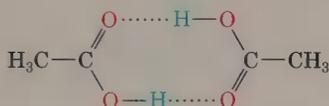
Since carboxylic acid functional groups are structurally related to both ketones and alcohols, we might expect to see some familiar properties. Carboxylic acids are indeed similar to both ketones and alcohols in some ways, though there are also major differences. Like ketones, the carboxyl carbon is  $sp^2$  hybridized. Carboxylic acid groups are therefore planar, with C–C–O and O–C–O bond angles of approximately  $120^\circ$ . The physical parameters of acetic acid are given in Table 20.2.



Like alcohols, carboxylic acids are strongly associated because of hydrogen bonding. Most carboxylic acids exist as dimers held together by two hydrogen bonds:

**TABLE 20.2** Physical Parameters for Acetic Acid

<i>Bond angle (degrees)</i>		<i>Bond length (Å)</i>	
C—C=O	119	C—C	1.52
C—C—OH	119	C=O	1.25
O=C—OH	122	C—OH	1.31



Acetic acid dimer

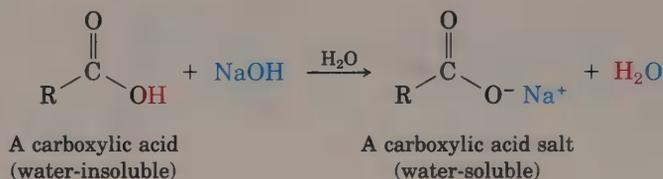
This strong hydrogen bonding has a noticeable effect on boiling points, making carboxylic acids much higher boiling than the corresponding alcohols. Table 20.3 lists the properties of some common acids.

**TABLE 20.3** Physical Constants of Some Carboxylic Acids

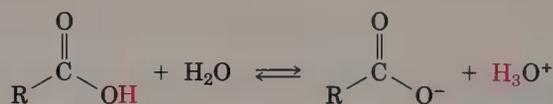
<i>Structure</i>	<i>Name</i>	<i>Melting point (°C)</i>	<i>Boiling point (°C)</i>
HCOOH	Formic	8.4	100.5
CH <sub>3</sub> COOH	Acetic	16.6	118
CH <sub>3</sub> CH <sub>2</sub> COOH	Propanoic	-22	141
CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	Butanoic	-4.2	163
FCH <sub>2</sub> COOH	Fluoroacetic	35.2	165
BrCH <sub>2</sub> COOH	Bromoacetic	50	208
HOCH <sub>2</sub> COOH	Glycolic	80	Decomposes
H <sub>2</sub> C=CHCOOH	Propenoic	13	141
C <sub>6</sub> H <sub>5</sub> COOH	Benzoic	122.4	249
HOCCOOH	Oxalic	189.5	Decomposes
HOOCCH <sub>2</sub> COOH	Malonic	135	Decomposes
HOOCCH <sub>2</sub> CH <sub>2</sub> COOH	Succinic	188	Decomposes
(Z)-HOOCCH=CHCOOH	Maleic	139	Decomposes

## 20.3 Dissociation of Carboxylic Acids

As their name implies, carboxylic acids are *acidic*. They therefore react with bases such as sodium hydroxide and sodium bicarbonate to give metal carboxylate salts. Although carboxylic acids with more than six carbons are only slightly soluble in water, alkali metal salts of carboxylic acids are generally quite water-soluble because of their ionic structure. It's often possible to take advantage of this solubility to purify acids by extracting their salts into aqueous base, then reacidifying and extracting the pure acid back into an organic solvent.



Like other Brønsted–Lowry acids discussed in Section 2.6, carboxylic acids dissociate slightly in dilute aqueous solution to give  $\text{H}_3\text{O}^+$  and carboxylate anion,  $\text{RCOO}^-$ .

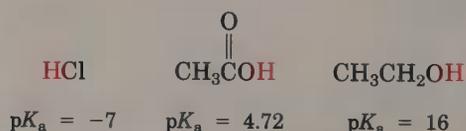


As with all acids, we can define an acidity constant,  $K_a$ :

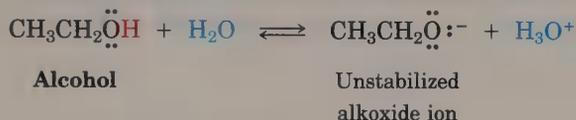
$$K_a = \frac{[\text{RCOO}^-][\text{H}_3\text{O}^+]}{[\text{RCOOH}]} \quad \text{and} \quad \text{p}K_a = -\log K_a$$

For most carboxylic acids, the acidity constant  $K_a$  is on the order of  $10^{-5}$ . Acetic acid, for example, has  $K_a = 1.8 \times 10^{-5}$ , which corresponds to a  $\text{p}K_a$  of 4.72. In practical terms,  $K_a$  values near  $10^{-5}$  mean that only about 1% of the molecules in a 0.1 M solution are dissociated, as opposed to the 100% dissociation found with strong mineral acids such as HCl and  $\text{H}_2\text{SO}_4$ .

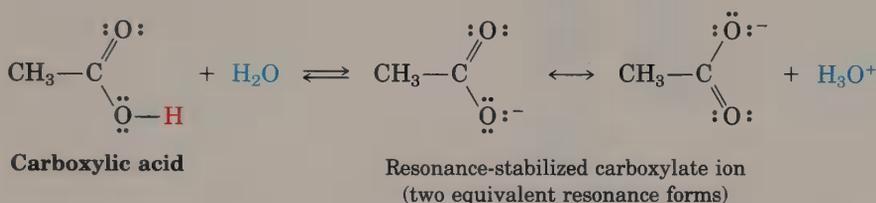
Although much weaker than mineral acids, carboxylic acids are nevertheless much stronger acids than alcohols are. The  $K_a$  of ethanol, for example, is approximately  $10^{-16}$ , making ethanol a weaker acid than acetic acid by a factor of  $10^{11}$ .



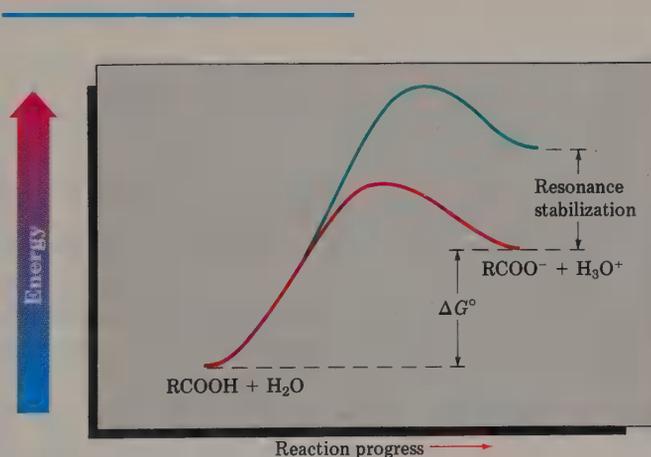
Why are carboxylic acids so much more acidic than alcohols even though both contain O–H groups? The easiest way to answer this question is to look at the relative stability of carboxylate anions versus alkoxide anions. Alkoxides are oxygen anions in which the negative charge is localized on a single electronegative atom:



Carboxylates are also oxygen anions, but the negative charge is *delocalized* over both oxygen atoms, resulting in stabilization of the ion. In resonance terms (Section 10.7), a carboxylate ion is a stabilized resonance hybrid of two equivalent Kekulé structures:

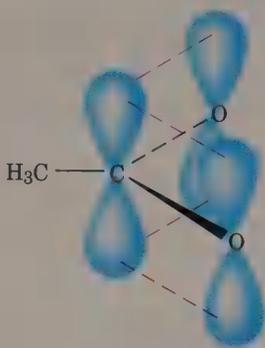


Since a carboxylate ion is more stable than an alkoxide ion, it is lower in energy and is present in greater amount at equilibrium, as shown on the reaction energy diagrams in Figure 20.1. Dissociation of a carboxylic acid has a smaller  $\Delta G^\circ$  than dissociation of an alcohol has, leading to a larger equilibrium constant,  $K_a$ .



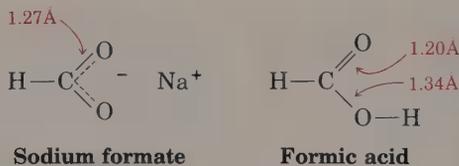
**Figure 20.1** A reaction energy diagram for the dissociation of an alcohol (green curve) and a carboxylic acid (red curve). Resonance stabilization of the carboxylate anion lowers  $\Delta G^\circ$  for dissociation of the acid, leading to a larger  $K_a$ . (The starting energy levels of alcohol and acid are shown at the same point for ease of comparison.)

We can't really draw an accurate representation of the carboxylate resonance hybrid using Kekulé structures, but an orbital picture of acetate ion is helpful in making it clear that the carbon–oxygen bonds are equivalent and that each is intermediate between single and double bonds (Figure 20.2). The  $p$  orbital on the carboxylate carbon atom overlaps equally well with  $p$  orbitals from *both* oxygens, and the four  $p$  electrons are delocalized throughout the three-atom pi electron system.



**Figure 20.2** An orbital picture of acetate ion, showing the equivalence of the two oxygen atoms.

Evidence for the equivalence of the two carboxylate oxygens comes from X-ray studies on sodium formate. Both carbon–oxygen bonds are 1.27 Å in length, midway between the C=O double bond (1.20 Å) and C–O single bond (1.34 Å) of formic acid.



**PROBLEM** .....

- 20.3** Use the expression  $\Delta G^\circ = -2.303 RT \log K_a$  to calculate values of  $\Delta G^\circ$  for the dissociation of ethanol ( $\text{p}K_a = 16.0$ ) and acetic acid ( $\text{p}K_a = 4.72$ ) at 300 K (27°C). The gas constant  $R$  has the value 1.98 cal/mol · K.

**PROBLEM** .....

- 20.4** Assume that you had a mixture of naphthalene and benzoic acid that you wanted to separate. How might you take advantage of the acidity of one component in the mixture to effect a separation?

**PROBLEM** .....

- 20.5** The  $K_a$  for dichloroacetic acid is  $5.5 \times 10^{-2}$ . Approximately what percentage of the acid is dissociated in a 0.1 M aqueous solution?
- .....

## 20.4 Substituent Effects on Acidity

The  $pK_a$  values listed in Table 20.4 indicate a substantial difference in acidities for different carboxylic acids. For example, trifluoroacetic acid ( $K_a = 0.59$ ) is more than 32,000 times as strong as acetic acid ( $K_a = 1.8 \times 10^{-5}$ ). How can we account for such differences?

**TABLE 20.4** Acidity of Some Carboxylic Acids

Structure	$K_a$	$pK_a$
HCl (hydrochloric acid) <sup>a</sup>	( $10^7$ )	(-7)
F <sub>3</sub> CCOOH	0.59	0.23
Cl <sub>3</sub> CCOOH	0.23	0.64
Cl <sub>2</sub> CHCOOH	$5.5 \times 10^{-2}$	1.26
FCH <sub>2</sub> COOH	$2.6 \times 10^{-3}$	2.59
ClCH <sub>2</sub> COOH	$1.4 \times 10^{-3}$	2.85
BrCH <sub>2</sub> COOH	$1.3 \times 10^{-3}$	2.89
ICH <sub>2</sub> COOH	$7.5 \times 10^{-4}$	3.12
HCOOH	$1.77 \times 10^{-4}$	3.75
HOCH <sub>2</sub> COOH	$1.5 \times 10^{-4}$	3.83
ClCH <sub>2</sub> CH <sub>2</sub> COOH	$1.04 \times 10^{-4}$	3.98
C <sub>6</sub> H <sub>5</sub> COOH	$6.46 \times 10^{-5}$	4.19
H <sub>2</sub> C=CHCOOH	$5.6 \times 10^{-5}$	4.25
C <sub>6</sub> H <sub>5</sub> CH <sub>2</sub> COOH	$4.9 \times 10^{-5}$	4.31
CH <sub>3</sub> COOH	$1.8 \times 10^{-5}$	4.72
CH <sub>3</sub> CH <sub>2</sub> COOH	$1.34 \times 10^{-5}$	4.87
CH <sub>3</sub> CH <sub>2</sub> OH (ethanol) <sup>a</sup>	( $10^{-16}$ )	(16)

Stronger  
acid

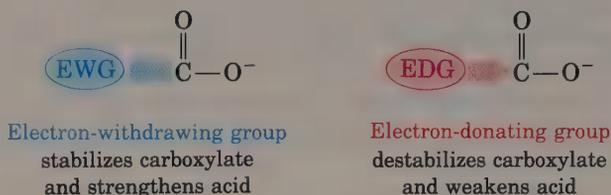


Weaker  
acid

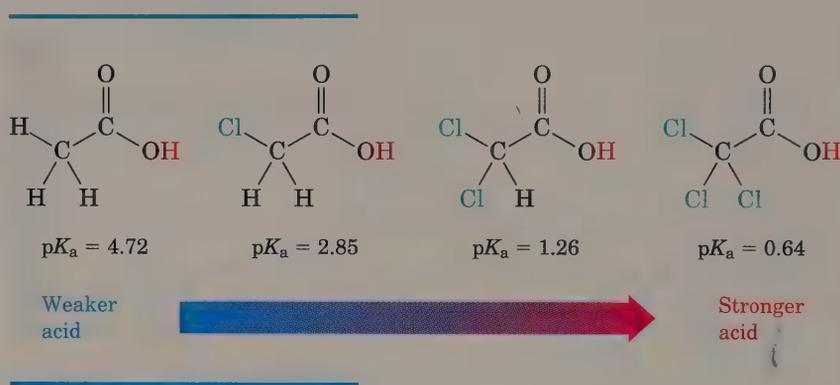
<sup>a</sup>Values for hydrochloric acid and ethanol are shown for reference.

Since the dissociation of a carboxylic acid is an equilibrium reaction, any factor that stabilizes the carboxylate anion product relative to undissociated carboxylic acid will drive the equilibrium toward increased dissociation and result in increased acidity. Conversely, any factor that destabilizes carboxylate relative to undissociated acid will result in decreased acidity. For example, an electron-*withdrawing* group attached to the carboxyl should inductively withdraw electron density, thereby delocalizing the negative charge, stabilizing the carboxylate anion, and increasing acidity. An electron-*donating* group, however, should have exactly the opposite effect

by concentrating the charge, destabilizing the carboxylate, and decreasing acidity.



The  $pK_a$  data in Table 20.4 show exactly the predicted effect. Highly electronegative substituents such as the halogens tend to make the carboxylate anion more stable by inductively withdrawing electrons. Thus fluoroacetic, chloroacetic, bromoacetic, and iodoacetic acids are all stronger than acetic acid by factors of 50–150. Introduction of two electronegative substituents makes dichloroacetic acid some 3000 times stronger than acetic acid, and introduction of three substituents makes trichloroacetic acid more than 12,000 times stronger (Figure 20.3).



**Figure 20.3** Relative strengths of chlorosubstituted acetic acids.

Since inductive effects are strongly dependent on distance, the effect of halogen substitution decreases as the substituent is moved farther from the carboxyl. The chlorobutanoic acids show clearly what happens as the electronegative substituent is moved successively farther from the carbonyl group (Table 20.5). 2-Chlorobutanoic acid has a  $pK_a$  of 2.86, the 3-substituted acid has a  $pK_a$  of 4.05, and the 4-substituted acid, with a  $pK_a$  of 4.52, has an acidity similar to that of butanoic acid itself.

**PROBLEM** .....

- 20.6** Rank the acids in the following groups in order of increasing acidity, without looking at a table of  $pK_a$  values:
- $\text{CH}_3\text{CH}_2\text{COOH}$ ,  $\text{BrCH}_2\text{COOH}$ ,  $\text{FCH}_2\text{COOH}$
  - Benzoic acid, *p*-nitrobenzoic acid, *p*-methoxybenzoic acid
  - $\text{CH}_3\text{CH}_2\text{OH}$ ,  $\text{CH}_3\text{CH}_2\text{NH}_2$ ,  $\text{CH}_3\text{CH}_2\text{COOH}$

TABLE 20.5 Acidity of Chlorosubstituted Butanoic Acids

Structure	$K_a$	$pK_a$
$\begin{array}{c} \text{Cl} \\   \\ \text{CH}_3\text{CH}_2\text{CHCOOH} \end{array}$	$1.39 \times 10^{-3}$	2.86
$\begin{array}{c} \text{Cl} \\   \\ \text{CH}_3\text{CHCH}_2\text{COOH} \end{array}$	$8.9 \times 10^{-5}$	4.05
$\text{ClCH}_2\text{CH}_2\text{CH}_2\text{COOH}$	$3 \times 10^{-5}$	4.52
$\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$	$1.5 \times 10^{-5}$	4.82

## PROBLEM

- 20.7 Dicarboxylic acids have two dissociation constants: one for the initial dissociation into a monoanion and one for the second dissociation into a dianion. For oxalic acid,  $\text{HOOC}-\text{COOH}$ , the first ionization constant has a  $pK_1$  of 1.2, and the second ionization constant has a  $pK_2$  of 4.2. Why is the second carboxyl group so much less acidic than the first?

## PROBLEM

- 20.8 Shown here are some  $pK_a$  data for simple dibasic acids. How do you account for the fact that the difference between the first and second ionization constants decreases with increasing distance between the carboxyl groups?

Name	Structure	$pK_1$	$pK_2$
Oxalic	$\text{HOCCOOH}$	1.2	4.2
Succinic	$\text{HOOCCH}_2\text{CH}_2\text{COOH}$	4.2	5.6
Adipic	$\text{HOOC}(\text{CH}_2)_4\text{COOH}$	4.4	4.4

## 20.5 Substituent Effects in Substituted Benzoic Acids

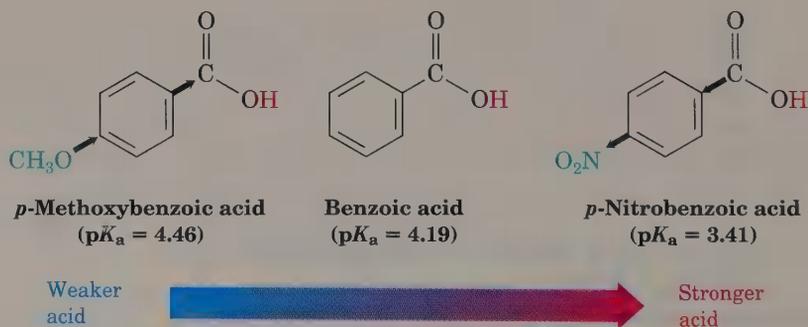
We saw during the discussion of electrophilic aromatic substitution (Section 16.5) that substituents on the aromatic ring play a large role in determining reactivity. Aromatic rings with electron-donating groups are activated toward further electrophilic substitution, and aromatic rings with electron-

withdrawing groups are deactivated. Exactly the same effects are noticed on the acidity of substituted benzoic acids (Table 20.6).

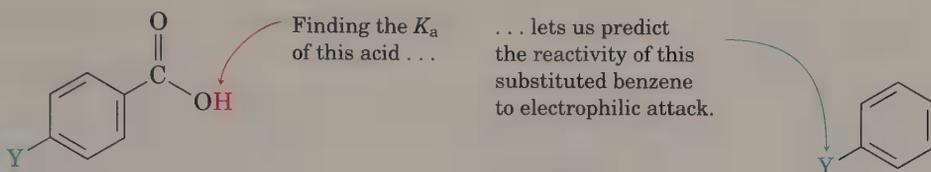
**TABLE 20.6** Substituent Effects on Acidity of Para-substituted Benzoic Acids

			
	Y	$K_a$	$pK_a$
Weaker acid	—OH	$2.8 \times 10^{-5}$	4.55
	—OCH <sub>3</sub>	$3.5 \times 10^{-5}$	4.46
	—CH <sub>3</sub>	$4.3 \times 10^{-5}$	4.34
Stronger acid	—H	$6.46 \times 10^{-5}$	4.19
	—Br	$1.1 \times 10^{-4}$	3.96
	—Cl	$1.1 \times 10^{-4}$	3.96
	—CHO	$1.8 \times 10^{-4}$	3.75
	—CN	$2.8 \times 10^{-4}$	3.55
	—NO <sub>2</sub>	$3.9 \times 10^{-4}$	3.41

As Table 20.6 shows, electron-withdrawing (deactivating) groups increase acidity by stabilizing the carboxylate anion, and electron-donating (activating) groups decrease acidity by destabilizing the carboxylate anion. Thus, an activating group such as methoxy decreases the acidity of benzoic acid, but a deactivating group such as nitro increases the acidity.



Since it's much easier to measure the acidity of a substituted benzoic acid than to determine the relative electrophilic reactivity of a substituted benzene, the correlation between the two effects can be valuable in predicting reactivity. If we want to know the effect of a certain substituent on electrophilic reactivity, we can simply find the acidity of the corresponding benzoic acid:



### PRACTICE PROBLEM.....

The  $pK_a$  of *p*-(trifluoromethyl)benzoic acid is 3.6. Would you expect the trifluoromethyl substituent to be an activating or deactivating group in the Friedel–Crafts reaction?

**Solution** A  $pK_a$  of 3.6 means that *p*-(trifluoromethyl)benzoic acid is stronger than benzoic acid, whose  $pK_a$  is 4.19. Thus, the trifluoromethyl substituent is favoring dissociation by helping to stabilize negative charge. Trifluoromethyl must therefore be an electron-withdrawing, deactivating group.

### PROBLEM.....

20.9 The  $pK_a$  of *p*-cyclopropylbenzoic acid is 4.45. Is cyclopropylbenzene likely to be more reactive or less reactive than benzene toward electrophilic bromination? Explain.

### PROBLEM.....

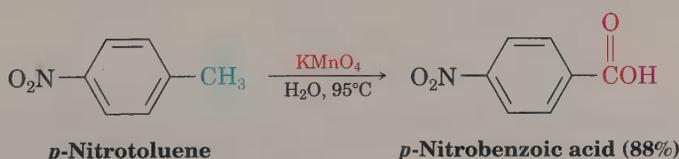
20.10 Rank the following compounds in order of increasing acidity. Don't look at a table of  $pK_a$  data to help with your answer.

- (a) Benzoic acid, *p*-methylbenzoic acid, *p*-chlorobenzoic acid  
 (b) *p*-Nitrobenzoic acid, acetic acid, benzoic acid

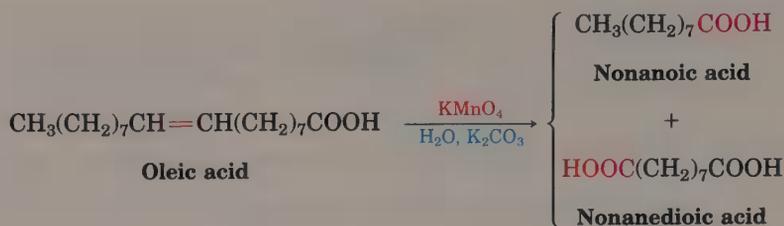
## 20.6 Preparation of Carboxylic Acids

We've already seen most of the common methods for preparing carboxylic acids, but let's review them briefly:

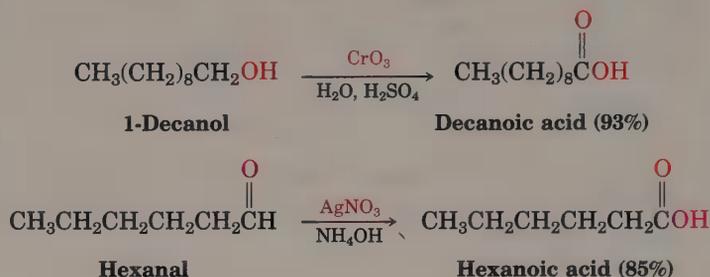
- Oxidation of substituted alkylbenzenes with potassium permanganate or sodium dichromate gives substituted benzoic acids (Section 16.11). Both primary and secondary alkyl groups can be oxidized in this manner, but tertiary groups are not affected.



- Oxidative cleavage of alkenes gives carboxylic acids if the alkene has at least one vinylic hydrogen (Section 7.9). The reaction can be carried out with sodium dichromate or potassium permanganate.

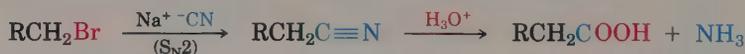


3. Oxidation of primary alcohols and aldehydes yields carboxylic acids (Sections 17.9 and 19.5). Primary alcohols are often oxidized with Jones' reagent ( $\text{CrO}_3$ ,  $\text{H}_2\text{O}$ ,  $\text{H}_2\text{SO}_4$ ), and aldehydes are oxidized with either Jones' reagent or basic silver oxide (Tollens' reagent). Both oxidations take place rapidly and in high yield.

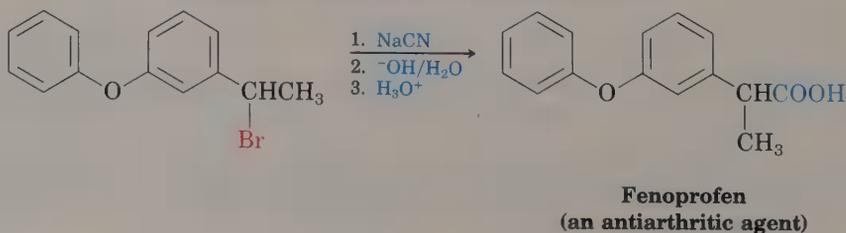


## HYDROLYSIS OF NITRILES

Nitriles,  $\text{R}-\text{C}\equiv\text{N}$ , can be hydrolyzed by strong aqueous acid or base to yield carboxylic acids. Since nitriles themselves are usually prepared by  $\text{S}_{\text{N}}2$  reaction between an alkyl halide and cyanide ion, the two-step sequence of cyanide displacement followed by nitrile hydrolysis is an excellent method for preparing carboxylic acids from alkyl halides ( $\text{RBr} \rightarrow \text{RC}\equiv\text{N} \rightarrow \text{RCOOH}$ ). Note that the product acid has one more carbon than the starting alkyl halide has.

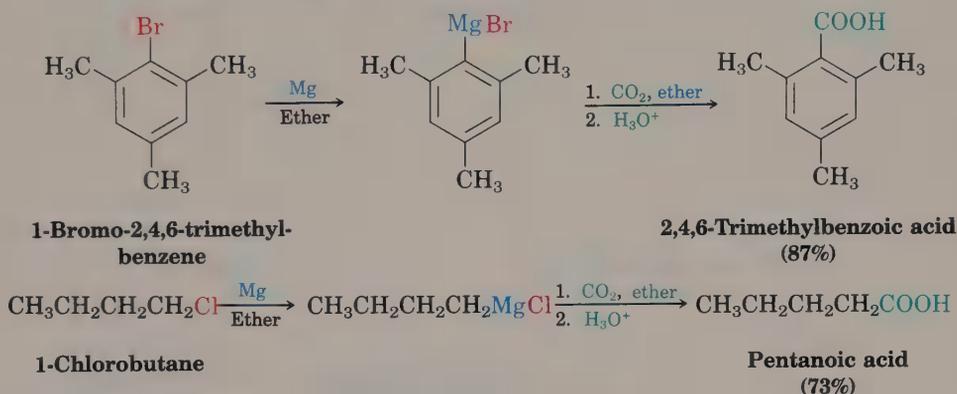


The method works best with primary halides, since a competitive  $\text{E}2$  elimination reaction can occur when a secondary or tertiary alkyl halide is used. Nevertheless, some unhindered secondary halides react well. A good example of the reaction occurs in the commercial synthesis of the antiarthritic drug fenoprofen.

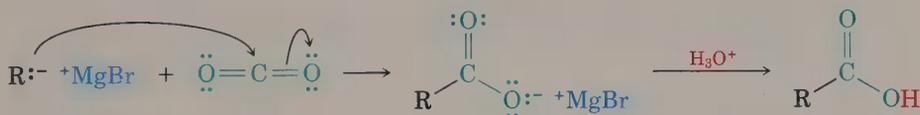


## CARBOXYLATION OF GRIGNARD REAGENTS

Yet another method for preparing carboxylic acids is by reaction of organo-metallic reagents with carbon dioxide. Both organomagnesium (Grignard) and organolithium reagents react with carbon dioxide to yield carboxylate salts, which can be protonated to give carboxylic acids. This **carboxylation** reaction is carried out either by pouring the Grignard reagent over dry ice (solid  $\text{CO}_2$ ) or by bubbling a stream of dry  $\text{CO}_2$  through the Grignard reagent solution. Grignard carboxylation generally gives good yields of acids from alkyl halides but is clearly limited in use to only those alkyl halides that can form Grignard reagents in the first place (Section 17.7).



The mechanism of Grignard carboxylation is similar to that of other Grignard reactions (Section 19.10). The organomagnesium halide adds to one of the  $\text{C}=\text{O}$  bonds of carbon dioxide in a typical nucleophilic addition reaction. Protonation of the carboxylate by addition of aqueous  $\text{HCl}$  then gives the free carboxylic acid product:



## PROBLEM.....

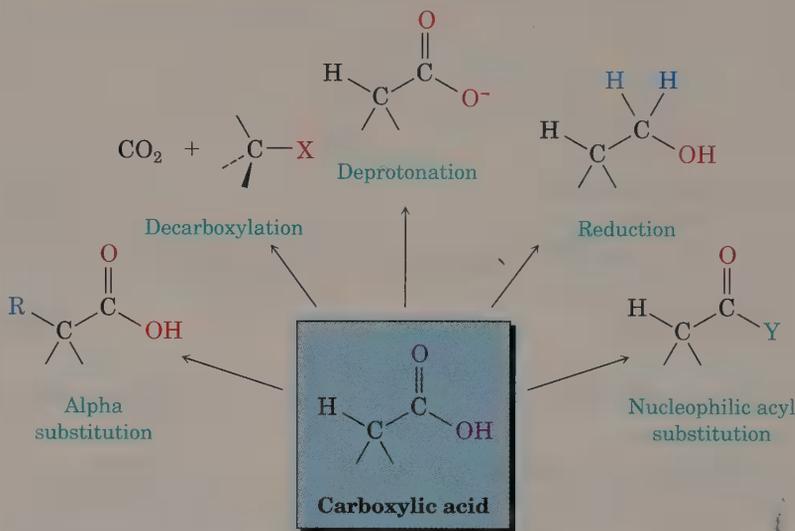
- 20.11** We've seen two methods for converting an alkyl halide into a carboxylic acid having one more carbon atom: (1) substitution with cyanide ion followed by hydrolysis and (2) formation of a Grignard reagent followed by carboxylation. What are the strengths and weaknesses of the two methods? Under what circumstances might one method be better than the other?

## PROBLEM.....

- 20.12** In light of your answer to Problem 20.11, what methods would you use to prepare the following carboxylic acids from organohalides?
- Benzoic acid from bromobenzene
  - $(\text{CH}_3)_3\text{CCOOH}$  from  $(\text{CH}_3)_3\text{CCl}$
  - $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$  from  $\text{CH}_3\text{CH}_2\text{CH}_2\text{Br}$

## 20.7 Reactions of Carboxylic Acids

We commented earlier in this chapter that carboxylic acids are structurally similar to both alcohols and ketones. As you might expect, there are also chemical similarities. Like alcohols, carboxylic acids can be deprotonated to give anions, which are good nucleophiles in  $S_N2$  reactions. Like ketones, carboxylic acids undergo attack by nucleophiles on their carbonyl group. In addition, carboxylic acids undergo other reactions characteristic neither of alcohols nor of ketones. Figure 20.4 shows some of the general types of reactions of carboxylic acids.

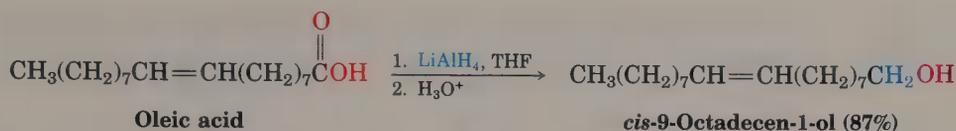


**Figure 20.4** Some general reactions of carboxylic acids.

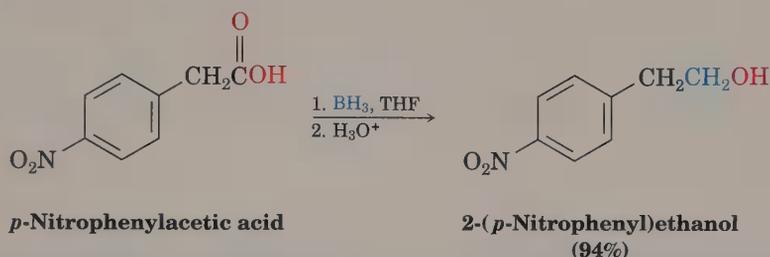
Reactions of carboxylic acids can be grouped into the five categories indicated in Figure 20.4. Of the five, we've already discussed the acidic behavior of carboxylic acids in Sections 20.3–20.5, and we'll discuss the decarboxylation reaction and reduction reaction later in this chapter. The remaining two general reactions of carboxylic acids—nucleophilic acyl substitution and alpha substitution—are two of the four fundamental carbonyl-group reaction mechanisms and will be discussed in detail in Chapters 21 and 22.

## 20.8 Reduction of Carboxylic Acids

Carboxylic acids are reduced by powerful hydride reagents such as lithium aluminum hydride (but not NaBH<sub>4</sub>) to yield primary alcohols (Section 17.6). The reaction is difficult, however, and often requires heating in tetrahydrofuran solvent to go to completion.

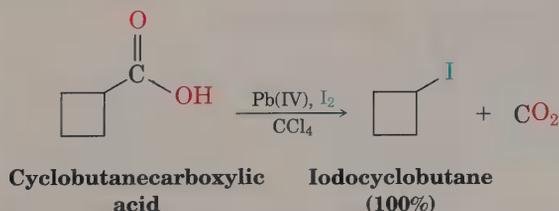
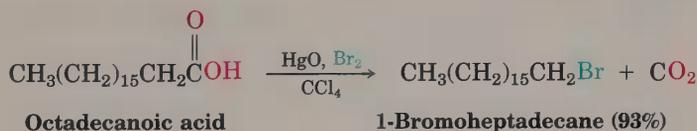


Borane,  $\text{BH}_3$ , is also used for converting carboxylic acids into primary alcohols. Reaction of an acid with borane occurs rapidly at room temperature, and the procedure is often preferred to reduction with  $\text{LiAlH}_4$  because of its relative ease, safety, and specificity. Borane reacts with carboxylic acids faster than with any other functional group, thereby allowing selective transformations such as that shown below on *p*-nitrophenylacetic acid. Were the reduction of *p*-nitrophenylacetic acid tried with  $\text{LiAlH}_4$ , both nitro and carboxyl groups would be reduced.



## 20.9 Decarboxylation of Carboxylic Acids: The Hunsdiecker Reaction

Carboxylic acids undergo **decarboxylation**—the loss of carbon dioxide—under certain conditions to give a product having one less carbon atom than the starting acid has. In the **Hunsdiecker<sup>1,2</sup> reaction**, which involves heating a heavy-metal salt of a carboxylic acid with bromine or iodine, carbon dioxide is lost and an alkyl halide having one less carbon atom than the starting acid is formed. The metal may be silver ion, mercuric ion, or lead(IV); all work equally well.



<sup>1</sup>Heinz Hunsdiecker (1904– ); b. Köln, Germany; Ph.D. Köln (Wintgen); private laboratory, Köln, Germany.

<sup>2</sup>Clare Hunsdiecker (1903– ); b. Kiel; Ph.D. Köln (Wintgen); private laboratory, Köln, Germany.

The Hunsdiecker reaction occurs by the radical chain pathway shown in Figure 20.5. The heavy-metal salt of the carboxylic acid reacts with bromine to form an acyl hypobromite,  $\text{RCOOBr}$ , a small amount of which undergoes homolytic cleavage of its weak  $\text{O}-\text{Br}$  bond to yield a carboxyl radical in the initiation step. The carboxyl radical then loses carbon dioxide in the first of two propagation steps to form an alkyl radical that abstracts a bromine from a second molecule of acyl hypobromite. This second propagation step yields the product bromoalkane and produces another alkyl radical to carry on the chain reaction.

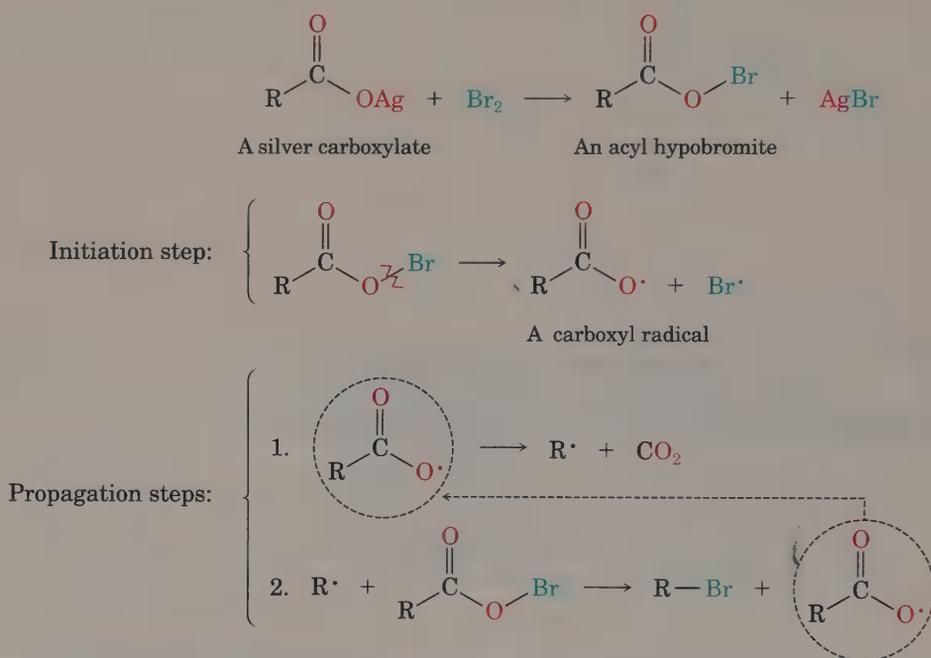


Figure 20.5 The mechanism of the Hunsdiecker reaction.

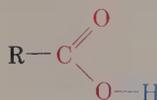
## 20.10 Spectroscopic Analysis of Carboxylic Acids

### INFRARED SPECTROSCOPY

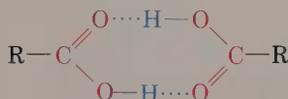
Carboxylic acids show two highly characteristic absorptions in the infrared spectrum that make this functional group easily identifiable. The  $\text{O}-\text{H}$  bond of the carboxyl group gives rise to a very broad absorption over the range  $2500\text{--}3300\text{ cm}^{-1}$ , and the  $\text{C}=\text{O}$  bond shows an absorption between  $1710\text{ cm}^{-1}$  and  $1760\text{ cm}^{-1}$ . The exact position of carbonyl absorption depends both on the structure of the molecule and on whether the acid is free (monomeric)

or associated (dimeric). Free carboxyl groups absorb at  $1760\text{ cm}^{-1}$ , but the more commonly encountered associated carboxyl groups absorb in a broad band centered around  $1710\text{ cm}^{-1}$ .

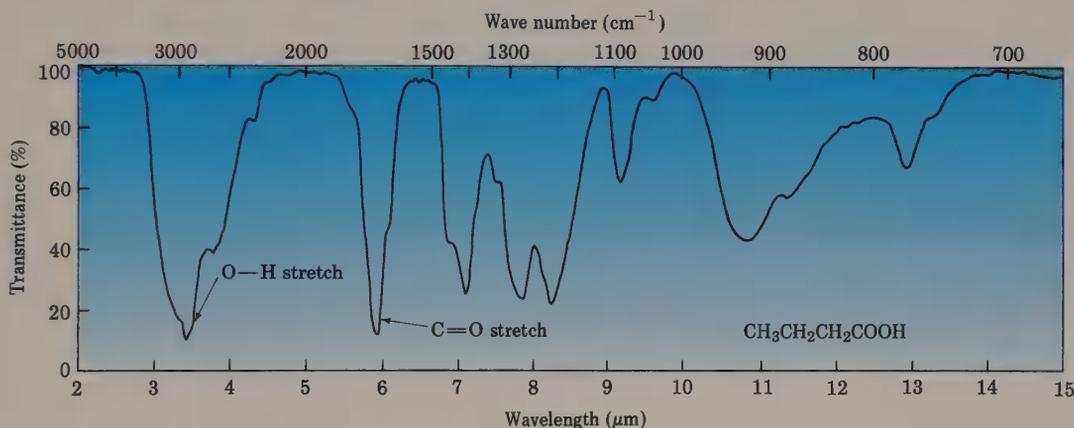
Free carboxyl  
(uncommon),  $1760\text{ cm}^{-1}$



Associated carboxyl  
(usual case),  $1710\text{ cm}^{-1}$



The infrared spectrum of butanoic acid shown in Figure 20.6 has both the broad O–H absorption and the C=O absorption at  $1710\text{ cm}^{-1}$  (associated) identified.

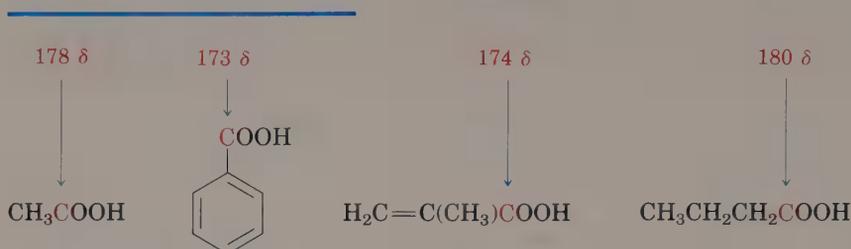


**Figure 20.6** Infrared spectrum of butanoic acid,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOH}$ .

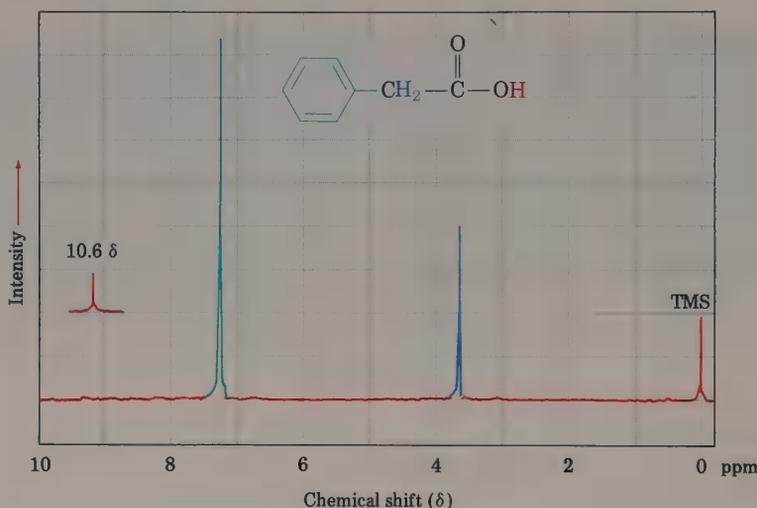
## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Carboxylic acid groups can be detected by both  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectroscopy. Carboxyl carbon atoms absorb in the range  $165\text{--}185\ \delta$  in the  $^{13}\text{C}$  NMR spectrum, with aromatic and  $\alpha,\beta$ -unsaturated acids near the upfield end of the range ( $\sim 165\ \delta$ ) and saturated aliphatic acids near the downfield end ( $\sim 185\ \delta$ ). The acidic  $-\text{COOH}$  proton normally absorbs near  $12\ \delta$  in the  $^1\text{H}$  NMR spectrum. Since the carboxylic acid proton has no neighbor protons, it is unsplit and occurs as a singlet. As with alcohols (Section 17.11), the  $-\text{COOH}$  proton can be replaced by deuterium when  $\text{D}_2\text{O}$  is added to the sample tube, causing the  $-\text{COOH}$  absorption to disappear from the NMR spectrum.

Figure 20.7 indicates the positions of the  $^{13}\text{C}$  NMR absorptions for several carboxylic acids, and Figure 20.8 shows the  $^1\text{H}$  NMR spectrum of phenylacetic acid. Note that the carboxyl proton occurs at  $10.6\ \delta$  and that all five aromatic ring protons show an accidental overlap as an apparent singlet at  $7.3\ \delta$ .



**Figure 20.7** Carbon-13 NMR absorptions of some carboxylic acids.

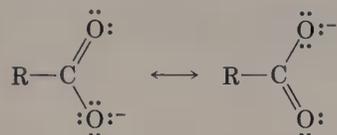


**Figure 20.8** Proton NMR spectrum of phenylacetic acid.

## 20.11 Summary and Key Words

**Carboxylic acids** are among the most important building blocks for synthesizing other molecules, both in nature and in the chemical laboratory. They are named systematically by replacing the terminal *-e* of the corresponding alkane name with *-oic acid*. Like ketones and aldehydes, the carbonyl carbon atom is  $sp^2$  hybridized; like alcohols, carboxylic acids are associated via hydrogen bonding and therefore have high boiling points.

The distinguishing characteristic of carboxylic acids is their acidity. Although weaker than mineral acids such as HCl, carboxylic acids nevertheless dissociate far more readily than alcohols. The reason for this difference lies in the stability of carboxylate ions: Carboxylate ions are stabilized by resonance between two equivalent forms:



Most alkanolic acids have  $\text{p}K_{\text{a}}$  values near 5, but the exact acidity constant of a given acid varies depending on structure. Carboxylic acids substituted by electron-withdrawing groups are more acidic (have a lower  $\text{p}K_{\text{a}}$ ) because their carboxylate ions are stabilized. Carboxylic acids substituted by electron-donating groups are less acidic (have a higher  $\text{p}K_{\text{a}}$ ) because their carboxylate ions are destabilized.

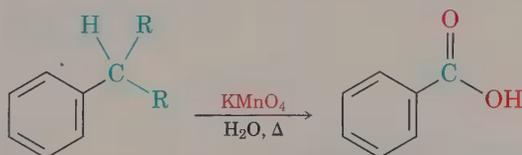
Methods of synthesis for carboxylic acids include: (1) oxidation of alkylbenzenes with sodium dichromate or potassium permanganate; (2) oxidative cleavage of alkenes; (3) oxidation of primary alcohols or aldehydes with Jones' reagent; (4) hydrolysis of nitriles; and (5) reaction of Grignard reagents with  $\text{CO}_2$  (**carboxylation**). The last two methods are particularly useful because they allow alkyl halides to be transformed into carboxylic acids with the addition of one carbon atom,  $\text{RBr} \rightarrow \text{RCOOH}$ .

General reactions of carboxylic acids include: (1) loss of the acidic proton; (2) nucleophilic acyl substitution at the carbonyl group; (3) loss of  $\text{CO}_2$  by a **decarboxylation** reaction; (4) substitution on the alpha carbon; and (5) reduction.

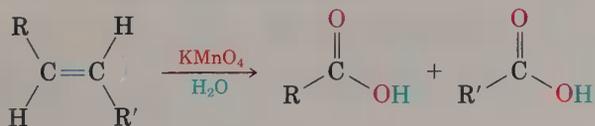
Carboxylic acids are easily distinguished spectroscopically. They show characteristic infrared absorptions at  $2500\text{--}3300\text{ cm}^{-1}$  (due to the O–H) and at  $1710\text{--}1760\text{ cm}^{-1}$  (due to the C=O). Acids also show  $^{13}\text{C}$  NMR absorptions at  $165\text{--}185\text{ }\delta$  and  $^1\text{H}$  NMR absorptions near  $12\text{ }\delta$ .

## 20.12 Summary of Reactions

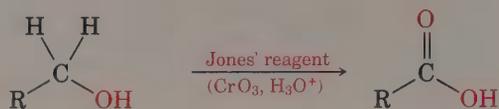
1. Preparation of carboxylic acids (Section 20.6)
  - a. Oxidation of alkylbenzenes (Section 16.11)



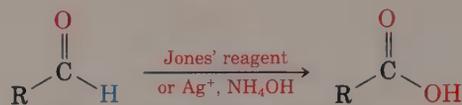
- b. Oxidative cleavage of alkenes (Section 7.9)



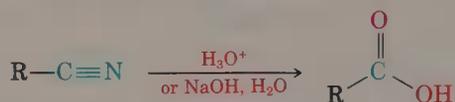
## c. Oxidation of primary alcohols (Section 17.9)



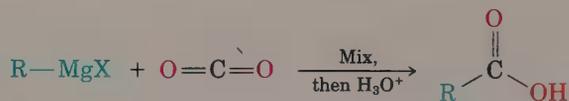
## d. Oxidation of aldehydes (Section 19.5)



## e. Hydrolysis of nitriles (Section 20.6)

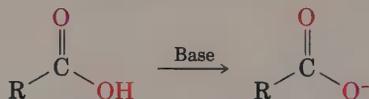


## f. Carboxylation of Grignard reagents (Section 20.6)

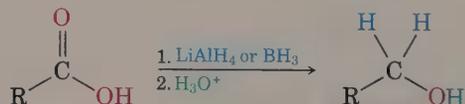


## 2. Reactions of carboxylic acids

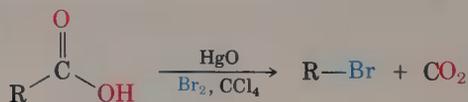
## a. Deprotonation (Section 20.3)



## b. Reduction to primary alcohols (Section 20.8)

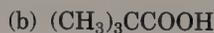
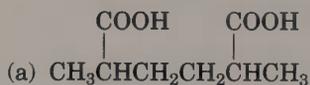


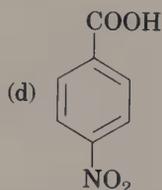
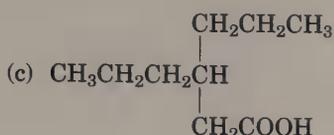
## c. Decarboxylation: Hunsdiecker reaction (Section 20.9)



## ADDITIONAL PROBLEMS .....

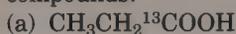
20.13 Give IUPAC names for these compounds:





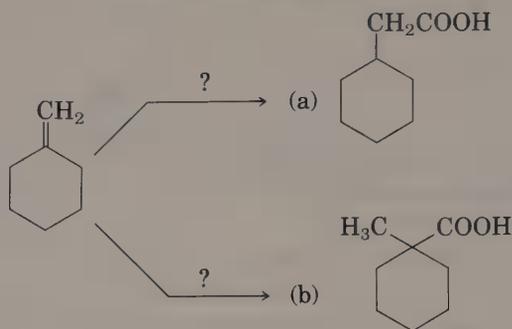
- 20.14** Draw structures corresponding to these IUPAC names:
- (a) *cis*-1,2-Cyclohexanedicarboxylic acid      (b) Heptanedioic acid  
 (c) 2-Hexen-4-ynoic acid      (d) 4-Ethyl-2-propyloctanoic acid  
 (e) 3-Chlorophthalic acid      (f) Triphenylacetic acid
- 20.15** Acetic acid boils at 118°C, but its ethyl ester boils at 77°C. Why is the boiling point of the acid so much higher, even though it has the lower molecular weight?
- 20.16** Draw and name the eight carboxylic acid isomers having the formula  $\text{C}_6\text{H}_{12}\text{O}_2$ .
- 20.17** Order the compounds in each set with respect to increasing acidity:
- (a) Acetic acid, oxalic acid, formic acid  
 (b) *p*-Bromobenzoic acid, *p*-nitrobenzoic acid, 2,4-dinitrobenzoic acid  
 (c) Phenylacetic acid, diphenylacetic acid, 3-phenylpropanoic acid  
 (d) Fluoroacetic acid, 3-fluoropropanoic acid, iodoacetic acid
- 20.18** Arrange the compounds in each set in order of increasing basicity:
- (a) Magnesium acetate, magnesium hydroxide, methylmagnesium bromide  
 (b) Sodium benzoate, sodium *p*-nitrobenzoate, sodium acetylide  
 (c) Lithium hydroxide, lithium ethoxide, lithium formate
- 20.19** Account for the fact that phthalic acid (1,2-benzenedicarboxylic acid) has  $\text{p}K_2 = 5.4$  but terephthalic acid (1,4-benzenedicarboxylic acid) has  $\text{p}K_2 = 4.8$ .
- 20.20** How could you convert butanoic acid into the following compounds? Write out each step showing the reagents needed.
- (a) 1-Butanol      (b) 1-Bromobutane      (c) Pentanoic acid  
 (d) 1-Butene      (e) 1-Bromopropane      (f) Octane
- 20.21** How could you convert each of the following compounds into butanoic acid? Write out each step showing all reagents.
- (a) 1-Butanol      (b) 1-Bromobutane      (c) 1-Butene  
 (d) 1-Bromopropane      (e) 4-Octene      (f) Pentanoic acid
- 20.22** How would you prepare these compounds from benzene? More than one step is required in each case.
- (a) *m*-Chlorobenzoic acid      (b) *p*-Bromobenzoic acid  
 (c) Phenylacetic acid,  $\text{C}_6\text{H}_5\text{CH}_2\text{COOH}$
- 20.23** Calculate  $\text{p}K_a$ 's for these acids:
- (a) Lactic acid,  $K_a = 8.4 \times 10^{-4}$       (b) Acrylic acid,  $\text{p}K_a = 5.6 \times 10^{-6}$
- 20.24** Calculate  $K_a$ 's for these acids:
- (a) Citric acid,  $\text{p}K_a = 3.14$       (b) Tartaric acid,  $\text{p}K_a = 2.98$
- 20.25** Predict the product of the reaction of *p*-methylbenzoic acid with each of the following reagents:
- (a)  $\text{BH}_3$ , then  $\text{H}_3\text{O}^+$       (b) *N*-Bromosuccinimide in  $\text{CCl}_4$   
 (c)  $\text{HgO}$ ,  $\text{I}_2$  in  $\text{CCl}_4$       (d)  $\text{CH}_3\text{MgBr}$  in ether, then  $\text{H}_3\text{O}^+$   
 (e)  $\text{KMnO}_4$ ,  $\text{H}_3\text{O}^+$       (f)  $\text{LiAlH}_4$ , then  $\text{H}_3\text{O}^+$

20.26 Using  $^{13}\text{C}$  as your only source of labeled carbon, how would you synthesize these compounds?

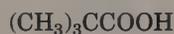
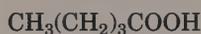


20.27 Propose a structure for an organic compound,  $\text{C}_6\text{H}_{12}\text{O}_2$ , that dissolves in dilute sodium hydroxide and that shows the following  $^1\text{H}$  NMR spectrum: 1.08  $\delta$  (9 H, singlet), 2.2  $\delta$  (2 H, singlet), and 11.2  $\delta$  (1 H, singlet).

20.28 How would you carry out these transformations?



20.29 What spectroscopic method could you use to distinguish among the following three isomeric acids? Tell exactly what characteristic features you would expect for each acid.

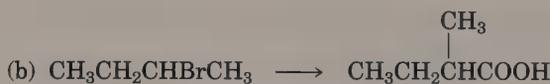


Pentanoic acid

3-Methylbutanoic acid

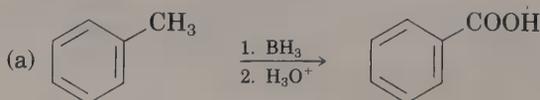
2,2-Dimethylpropanoic acid

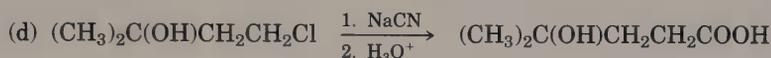
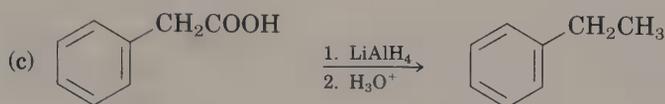
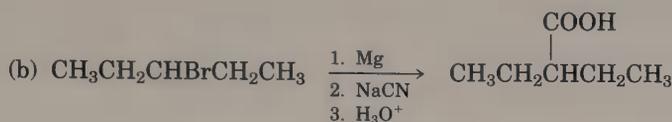
20.30 Which method of acid synthesis—Grignard carboxylation or nitrile hydrolysis—would you use for each of the following reactions? Explain your choices.



20.31 A chemist in need of 2,2-dimethylpentanoic acid decided to synthesize some by reaction of 2-chloro-2-methylpentane with  $\text{NaCN}$  followed by hydrolysis of the product. After carrying out the reaction sequence, however, none of the desired product could be found. What do you suppose went wrong?

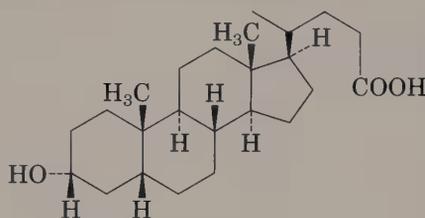
20.32 The following synthetic schemes all have at least one flaw in them. What is wrong with each?





**20.33** *p*-Aminobenzoic acid (PABA) is widely used as a sunscreen agent. Propose a synthesis of PABA starting from toluene.

**20.34** Lithocholic acid is a steroid found in human bile:

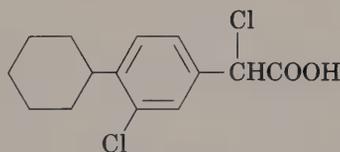


**Lithocholic acid**

Predict the product of reaction of lithocholic acid with each of the following reagents. Don't worry about the size of the molecule; just concentrate on the functional groups.

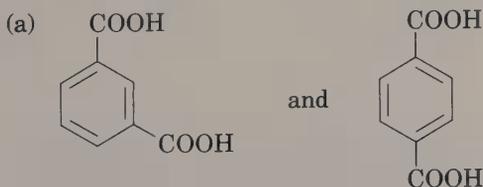
- |  |   |
|--|---|
| (a) Jones' reagent   | (b) Tollens' reagent  |
| (c) $\text{BH}_3$ , then $\text{H}_3\text{O}^+$            | (d) $(\text{CH}_3)_3\text{SiCl}$ , $(\text{CH}_3\text{CH}_2)_3\text{N}$ |
| (e) $\text{CH}_3\text{MgBr}$ , then $\text{H}_3\text{O}^+$ | (f) $\text{LiAlH}_4$ , then $\text{H}_3\text{O}^+$                      |

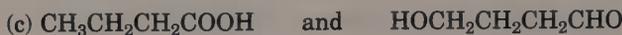
**20.35** Propose a synthesis of the anti-inflammatory drug Fenclorac from phenylcyclohexane.



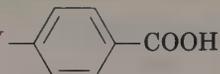
**Fenclorac**

**20.36** How would you use NMR (either  $^{13}\text{C}$  or  $^1\text{H}$ ) to distinguish between the following isomeric pairs?

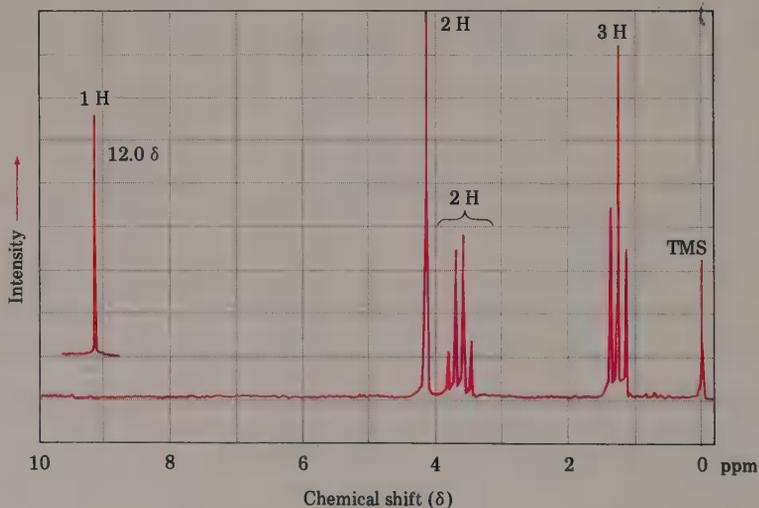




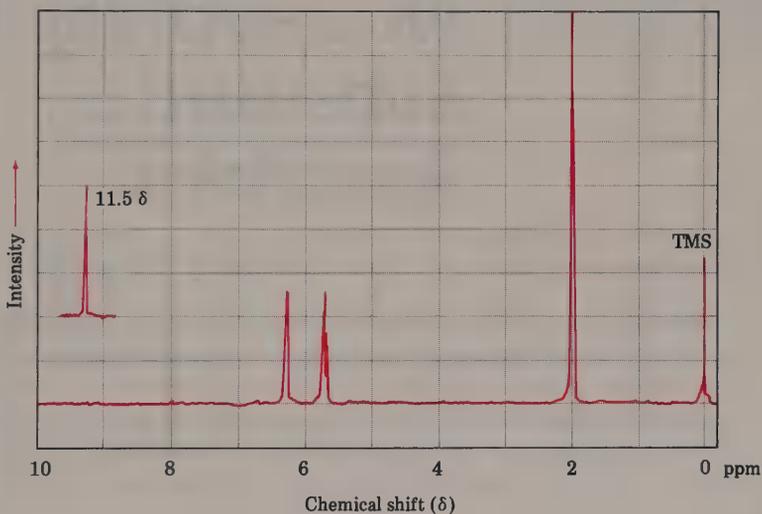
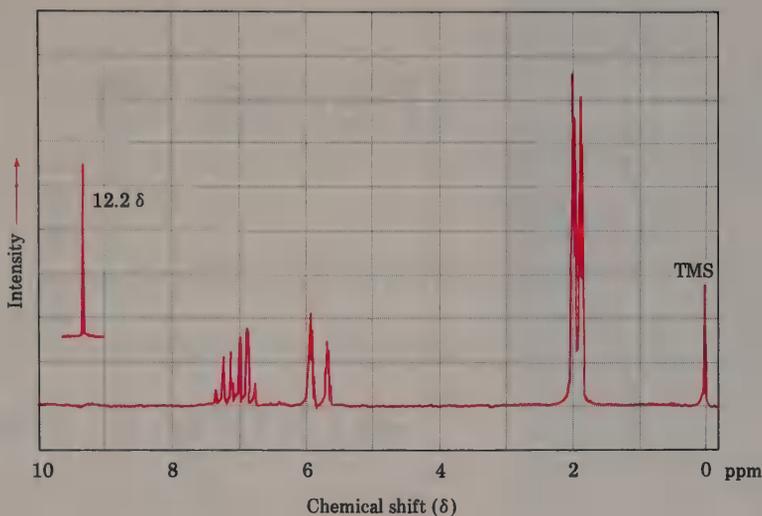
- 20.37** The  $\text{p}K_a$ 's of five para-substituted benzoic acids ( $\text{YC}_6\text{H}_4\text{COOH}$ ) are given in the following table. Rank the corresponding substituted benzenes ( $\text{YC}_6\text{H}_5$ ) in order of their increasing reactivity toward electrophilic aromatic substitution. If benzoic acid has  $\text{p}K_a = 4.20$ , which of the substituent groups are activators and which are deactivators?

Substituent Y	$\text{p}K_a$ of 
$-\text{Si}(\text{CH}_3)_3$	4.27
$-\text{CH}=\text{CHC}\equiv\text{N}$	4.03
$-\text{HgCH}_3$	4.10
$-\text{OSO}_2\text{CH}_3$	3.84
$-\text{PCl}_2$	3.59

- 20.38** Compound A,  $\text{C}_4\text{H}_8\text{O}_3$ , has infrared absorptions at  $1710$  and  $2500\text{--}3100\text{ cm}^{-1}$  and exhibits the  $^1\text{H}$  NMR spectrum shown. Propose a structure for A that is consistent with the data.



- 20.39** The two  $^1\text{H}$  NMR spectra shown here belong to crotonic acid (*trans*- $\text{CH}_3\text{CH}=\text{CHCOOH}$ ) and to methacrylic acid [ $\text{H}_2\text{C}=\text{C}(\text{CH}_3)\text{COOH}$ ]. Which spectrum corresponds to which acid? Explain.

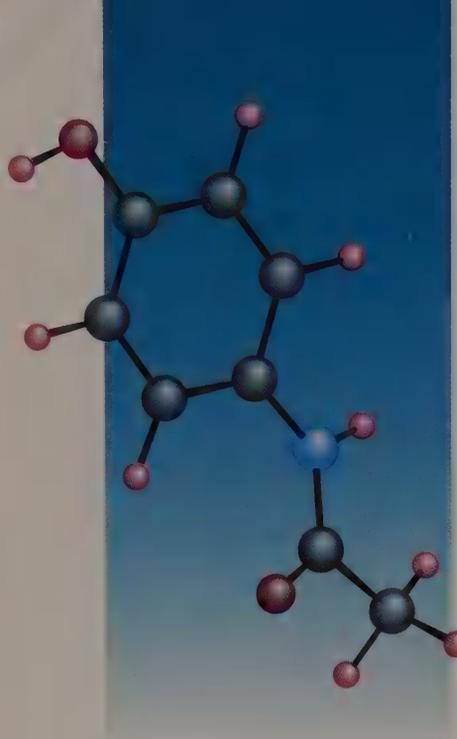


**20.40** Propose structures for carboxylic acids that show the following peaks in their proton-coupled  $^{13}\text{C}$  NMR spectra:

(a)  $\text{C}_7\text{H}_{12}\text{O}_2$ : 25.5  $\delta$  (t), 25.9  $\delta$  (t), 29.0  $\delta$  (t), 43.1  $\delta$  (d), 183.0  $\delta$  (s)

(b)  $\text{C}_8\text{H}_8\text{O}_2$ : 21.4  $\delta$  (q), 128.3  $\delta$  (s), 129.0  $\delta$  (d), 129.7  $\delta$  (d), 143.1  $\delta$  (s), 168.2  $\delta$  (s)

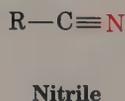
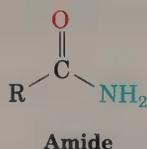
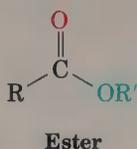
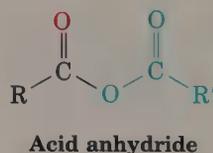
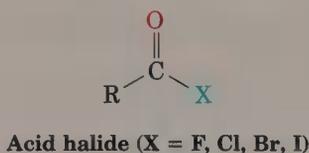
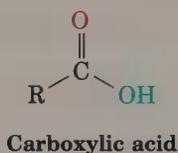
**20.41** 3-Methyl-2-hexenoic acid (mixture of *E* and *Z* isomers) has recently been identified as the substance responsible for the odor of human sweat. Devise a synthesis of the compound from starting materials having five or fewer carbons.



# Chapter 21

## Carboxylic Acid Derivatives and Nucleophilic Acyl Substitution Reactions

Carboxylic acids are just one member of a class of **acyl derivatives**,  $\text{RCOY}$ , where the acyl group is bonded to an electronegative substituent  $-\text{Y}$  that can act as a leaving group in substitution reactions. Numerous acyl derivatives are possible, but we'll be concerned only with four of the more common ones, in addition to carboxylic acids themselves: acid halides, acid anhydrides, esters, and amides. Also in this chapter, we'll discuss **nitriles**, a class of compounds closely related to carboxylic acids.



The chemistry of these acyl derivatives is similar and is dominated by a single general reaction: the **nucleophilic acyl substitution reaction** that we saw briefly in the overview of carbonyl chemistry:



Let's first learn more about acyl derivatives and then explore the chemistry of acyl substitution reactions.

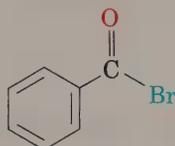
## 21.1 Nomenclature of Carboxylic Acid Derivatives

### ACID HALIDES: RCOX

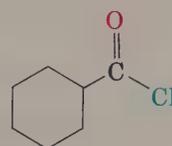
Acid halides are named by identifying first the acyl group and then the halide. The acyl group name is derived from the carboxylic acid name by replacing the *-ic acid* ending with *-yl* or the *-carboxylic acid* ending with *-carbonyl*. For example:



Acetyl chloride  
(from acetic acid)



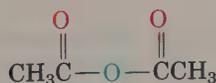
Benzoyl bromide  
(from benzoic acid)



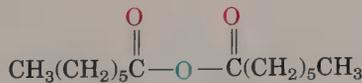
Cyclohexanecarbonyl chloride  
(from cyclohexanecarboxylic acid)

### ACID ANHYDRIDES: RCO<sub>2</sub>COR'

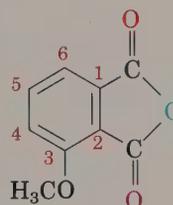
Symmetrical anhydrides of straight-chain monocarboxylic acids, and cyclic anhydrides of dicarboxylic acids, are named by replacing the word *acid* with *anhydride*:



Acetic anhydride

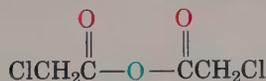


Heptanoic anhydride



3-Methoxyphthalic anhydride

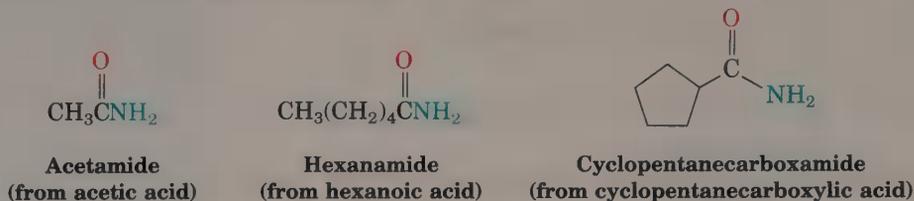
If the anhydride is derived from a substituted monocarboxylic acid, it is named by adding the prefix *bis-* (meaning two) to the acid name:



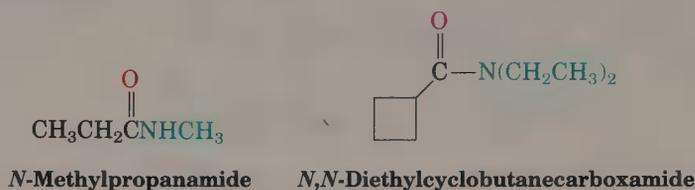
Bis(chloroacetic) anhydride

AMIDES:  $\text{RCONH}_2$ 

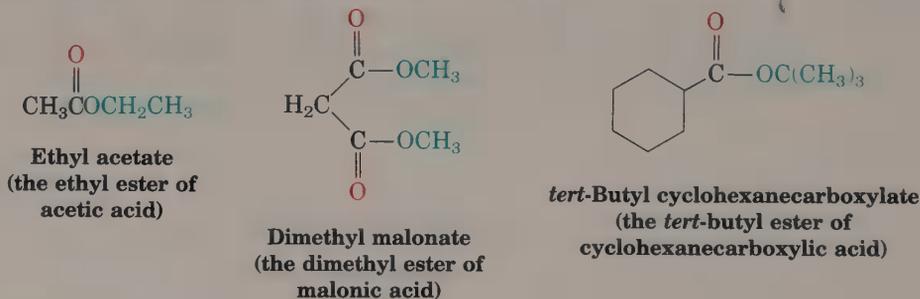
Amides with an unsubstituted  $-\text{NH}_2$  group are named by replacing the *-oic acid* or *-ic acid* ending with *-amide* or by replacing the *-carboxylic acid* ending with *-carboxamide*. For example:



If the nitrogen atom is further substituted, the compound is named by first identifying the substituent groups and then citing the parent amide name. The substituents are preceded by the letter *N* to identify them as being directly attached to nitrogen.

ESTERS:  $\text{RCO}_2\text{R}'$ 

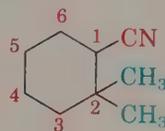
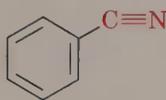
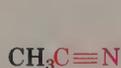
Systematic names for esters are derived by first giving the name of the alkyl group attached to oxygen and then identifying the carboxylic acid. In so doing, the *-ic acid* ending is replaced by *-ate*:

NITRILES:  $\text{RC}\equiv\text{N}$ 

Compounds containing the  $-\text{C}\equiv\text{N}$  functional group are known as **nitriles**. Simple acyclic nitriles are named by adding *-nitrile* as a suffix to the alkane name, with the nitrile carbon itself numbered C1.



More complex nitriles are named as derivatives of carboxylic acids by replacing the *-ic acid* or *-oic acid* ending with *-onitrile*, or by replacing the *-carboxylic acid* ending with *-carbonitrile*. In this system, the nitrile carbon atom is attached to C1 but is not itself numbered:



**Acetonitrile**  
(from acetic acid)

**Benzonitrile**  
(from benzoic acid)

**2,2-Dimethylcyclohexanecarbonitrile**  
(from 2,2-dimethylcyclohexane-  
carboxylic acid)

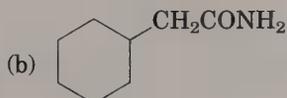
A summary of nomenclature rules for carboxylic acid derivatives is given in Table 21.1.

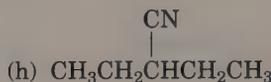
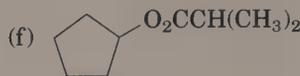
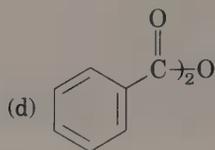
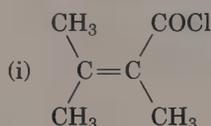
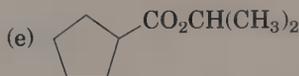
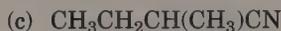
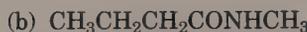
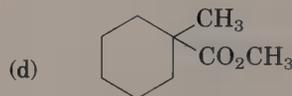
**TABLE 21.1** Nomenclature of Carboxylic Acid Derivatives

Functional group	Structure	Name ending
Carboxylic acid		<i>-ic acid</i> ( <i>-carboxylic acid</i> )
Acid halide		<i>-yl halide</i> ( <i>-carbonyl halide</i> )
Acid anhydride		<i>anhydride</i>
Amide		<i>-amide</i> ( <i>-carboxamide</i> )
Ester		<i>-ate</i> ( <i>-carboxylate</i> )
Nitrile	$\text{R}-\text{C}\equiv\text{N}$	<i>-onitrile</i> ( <i>-carbonitrile</i> )

**PROBLEM** .....

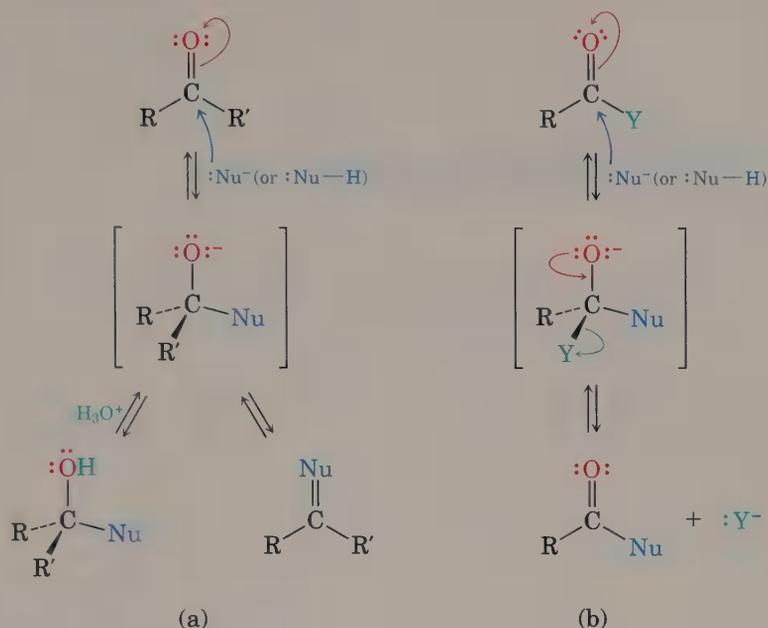
**21.1** Give IUPAC names for these structures:



**PROBLEM** .....**21.2** The names shown for the following compounds are incorrect. Give the correct names.**1-Pentenitrile****Methylbutanamide****2,4-Methylpentanoyl chloride****Methyl-2-methylcyclohexane carboxylate****21.2 Nucleophilic Acyl Substitution Reactions**

The addition of a nucleophile to the polar  $\text{C}=\text{O}$  bond is a general feature of most carbonyl-group reactions and is the first step in two of the four major carbonyl-group reactions. When nucleophiles add to aldehydes and ketones, the initially formed tetrahedral intermediate either can be protonated to yield an alcohol or can eliminate the carbonyl oxygen, leading to a new  $\text{C}=\text{Nu}$  bond. When nucleophiles add to carboxylic acid derivatives, however, a different reaction course is followed. The initially formed tetrahedral intermediate expels one of the two substituents originally bonded to the carbonyl carbon, leading to a net nucleophilic acyl substitution (Figure 21.1).

The different behavior of ketones/aldehydes and carboxylic acid derivatives is simply a consequence of structure. Carboxylic acid derivatives have an acyl function bonded to a potential leaving group,  $\text{Y}$ , that can leave as



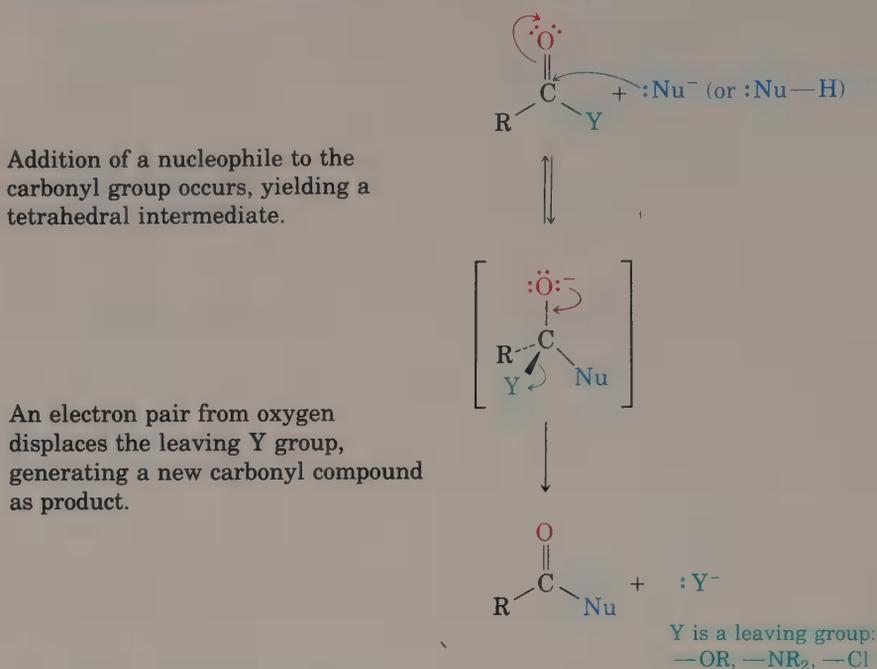
**Figure 21.1** The reactions of carbonyl groups with nucleophiles. (a) Ketones and aldehydes undergo nucleophilic addition reactions to yield either alcohols or products with a  $\text{C}=\text{Nu}$  double bond, whereas (b) carboxylic acid derivatives undergo nucleophilic acyl substitution reactions.

a stabilized anion. As soon as the tetrahedral intermediate is formed, the negative charge on oxygen expels this leaving group to generate a new carbonyl compound. Ketones and aldehydes have no such leaving group, however, and therefore don't undergo elimination.

As shown in Figure 21.2, the net effect of the two-step addition–elimination sequence is a substitution of the attacking nucleophile for the Y group originally bonded to the acyl carbon. Thus, the overall reaction is superficially similar to the kind of nucleophilic substitution that occurs during  $\text{S}_{\text{N}}2$  reactions (Section 11.5), but the mechanisms of the two reactions are completely different.  $\text{S}_{\text{N}}2$  reactions occur in a single step by back-side displacement of the leaving group, whereas nucleophilic acyl substitutions take place in two steps and involve a tetrahedral intermediate.

## 21.3 Relative Reactivity of Carboxylic Acid Derivatives

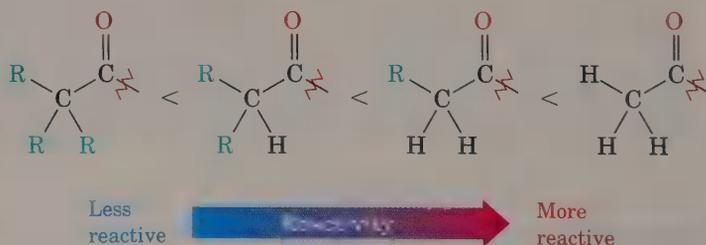
Nucleophilic acyl substitution reactions take place in two steps: addition of the nucleophile and elimination of a leaving group. Although both steps can sometimes affect the overall rate of reaction, it's generally the first step that



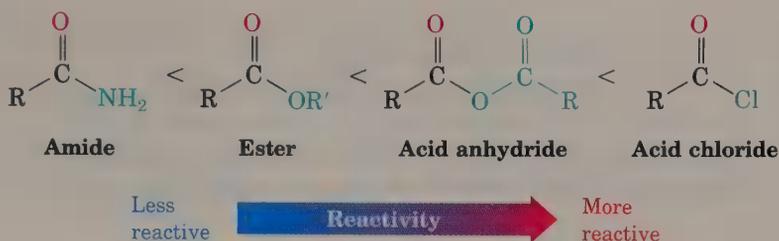
**Figure 21.2** General mechanism of nucleophilic acyl substitution.

is rate-limiting. Thus, any factor that makes the carbonyl group more easily attacked by nucleophiles favors the reaction.

Steric and electronic factors are both important in determining reactivity. Sterically, we find within a series of the same acid derivatives that unhindered, accessible carbonyl groups react with nucleophiles more readily than sterically hindered groups. For example, acetyl chloride,  $\text{CH}_3\text{COCl}$ , is much more reactive than 2,2-dimethylpropanoyl chloride,  $(\text{CH}_3)_3\text{CCOCl}$ . Thus, we find a reactivity order:

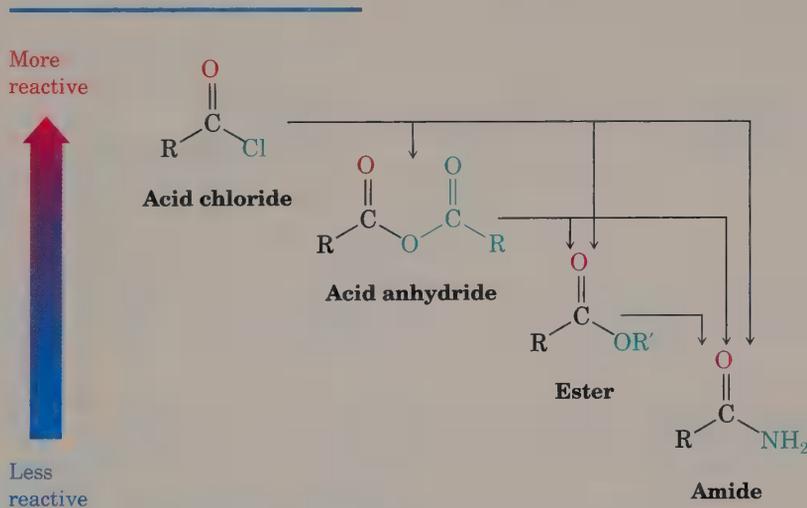


Electronically, we find that strongly polarized acyl derivatives are attacked more readily than less polar derivatives. Thus, acid chlorides are the most reactive acyl derivatives because the electronegative chlorine atom polarizes the carbonyl group more strongly than alkoxy or amino groups. The observed reactivity order is:



The way that various substituents affect the polarization of a carbonyl group is similar to the way they affect the reactivity of an aromatic ring toward electrophilic substitution (Section 16.5). Thus, a chlorine substituent withdraws electrons from an aromatic ring in the same way that it withdraws electrons from a neighboring acyl group, whereas amino and methoxyl substituents donate electrons to aromatic rings in the same way that they donate electrons to acyl groups.

An important consequence of the observed reactivity order is that *it is usually possible to transform a more reactive acid derivative into a less reactive one*. As we'll see in the next few sections, acid chlorides can be converted into anhydrides, esters, and amides, but amides can't readily be converted into esters, anhydrides, or acid chlorides. Remembering the reactivity order is therefore a way to keep track of a large number of reactions. Figure 21.3 shows the transformations that can be carried out.



**Figure 21.3** Interconversions of carboxylic acid derivatives.

A second consequence of the reactivity differences among acid derivatives is that only esters and amides are commonly found in nature. Acid halides and acid anhydrides don't occur naturally because they are too reactive. They rapidly undergo nucleophilic attack by water and are too reactive to exist in living organisms. Esters and amides, however, have exactly the

right balance of reactivity to allow them to occur widely and to be vitally important in many life processes.

In studying the chemistry of acid derivatives, we'll find that there are striking similarities among the various types of compounds. We'll be concerned largely with the reactions of just a few nucleophiles and will see that the same kinds of reactions keep occurring (Figure 21.4).

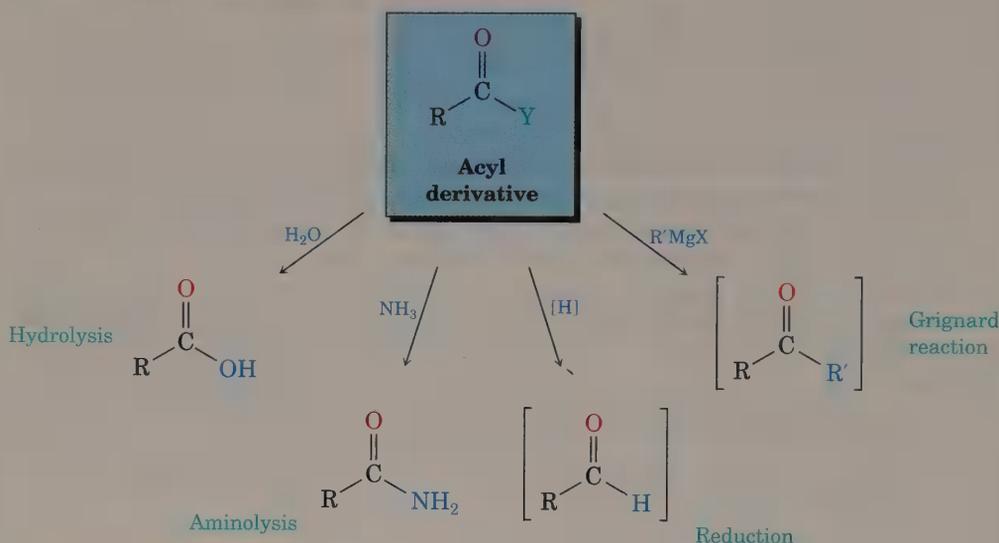
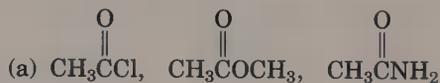


Figure 21.4 Some general reactions of carboxylic acid derivatives.

PROBLEM .....

- 21.3 Rank the compounds in the following sets with regard to expected reactivity toward nucleophilic acyl substitution:

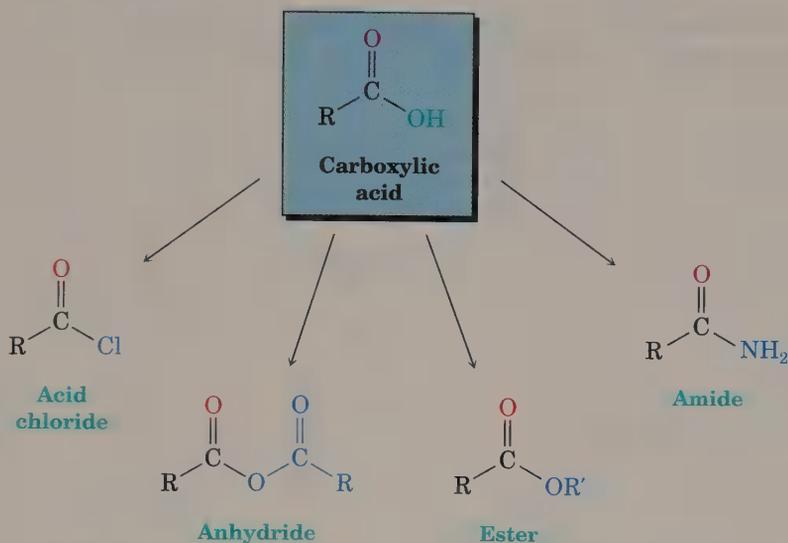


PROBLEM .....

- 21.4 How can you account for the fact that methyl trifluoroacetate,  $\text{CF}_3\text{COOCH}_3$ , is more reactive than methyl acetate,  $\text{CH}_3\text{COOCH}_3$ , in nucleophilic acyl substitution reactions?
- .....

## 21.4 Nucleophilic Acyl Substitution Reactions of Carboxylic Acids

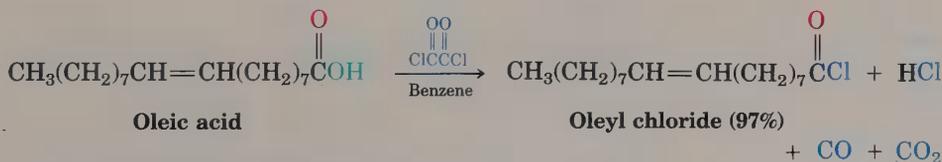
The most important reactions of carboxylic acids are those that convert the carboxyl group into other acid derivatives by nucleophilic acyl substitution,  $\text{RCOOH} \rightarrow \text{RCOY}$ . Acid chlorides, anhydrides, esters, and amides can all be prepared from carboxylic acids (Figure 21.5).

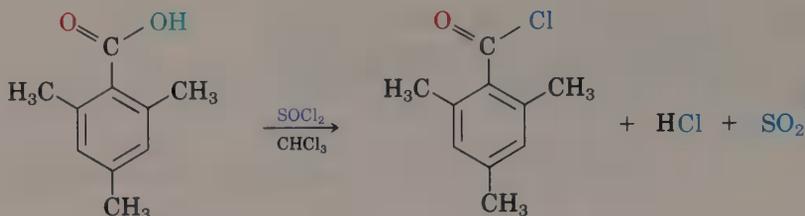


**Figure 21.5** Some nucleophilic acyl substitution reactions of carboxylic acids.

### CONVERSION INTO ACID CHLORIDES

Carboxylic acids are converted into carboxylic acid chlorides by treatment with thionyl chloride ( $\text{SOCl}_2$ ), phosphorus trichloride ( $\text{PCl}_3$ ), or oxalyl chloride ( $\text{ClCOCOCl}$ ). Thionyl chloride is both inexpensive and convenient to use but is strongly acidic; only acid-stable molecules can survive the reaction conditions. Oxalyl chloride, on the other hand, is much more expensive but gives higher yields and reacts under milder conditions.

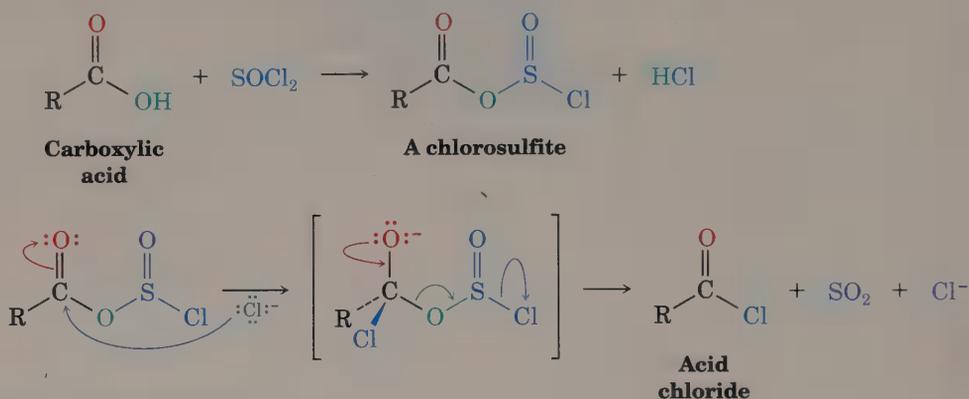




2,4,6-Trimethylbenzoic acid

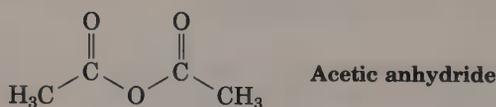
2,4,6-Trimethylbenzoyl  
chloride (90%)

These reactions occur by nucleophilic acyl substitution pathways in which the carboxylic acid is converted into a reactive derivative that is then attacked by a nucleophilic chloride ion. When  $\text{SOCl}_2$  is used, for example, the acid is first converted into a *chlorosulfite* intermediate:

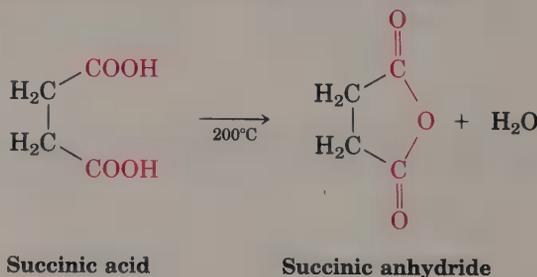


## CONVERSION INTO ACID ANHYDRIDES

Acid anhydrides are derived from two molecules of carboxylic acid by removing one equivalent of water. Acyclic anhydrides are difficult to prepare directly from the corresponding acids, and only acetic anhydride is commercially available.

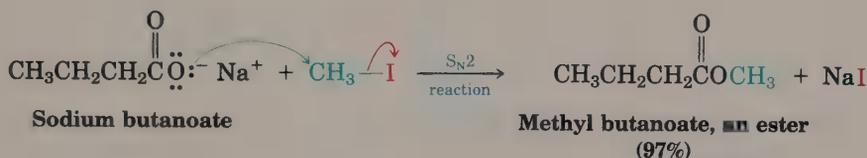


Cyclic anhydrides of ring size five or six are readily obtained by high-temperature dehydration of the diacids:

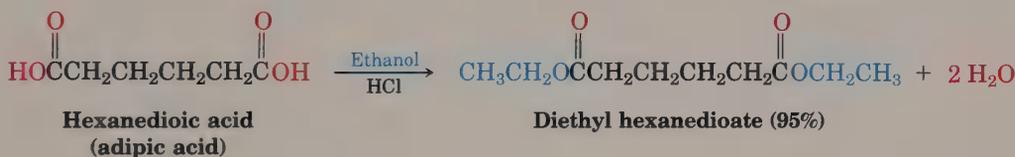
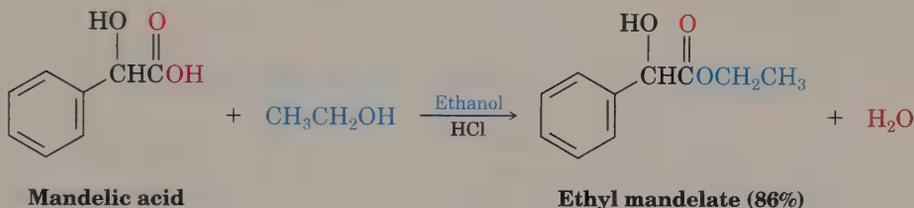


## CONVERSION INTO ESTERS

One of the most important reactions of carboxylic acids is their conversion into esters,  $\text{RCOOH} \rightarrow \text{RCOOR}'$ . There are many excellent methods for accomplishing this transformation, including the  $\text{S}_{\text{N}}2$  reaction between a carboxylate anion nucleophile and a primary alkyl halide that we've already studied (Section 11.5).



Esters can also be synthesized by a nucleophilic acyl substitution reaction between a carboxylic acid and an alcohol. Fischer<sup>1</sup> and Speier discovered in 1895 that esters result from simply heating a carboxylic acid in methanol or ethanol solution containing a small amount of mineral acid catalyst. Yields are good in this **Fischer esterification reaction**, but the need to use excess alcohol as solvent effectively limits the method to the synthesis of methyl, ethyl, and propyl esters.

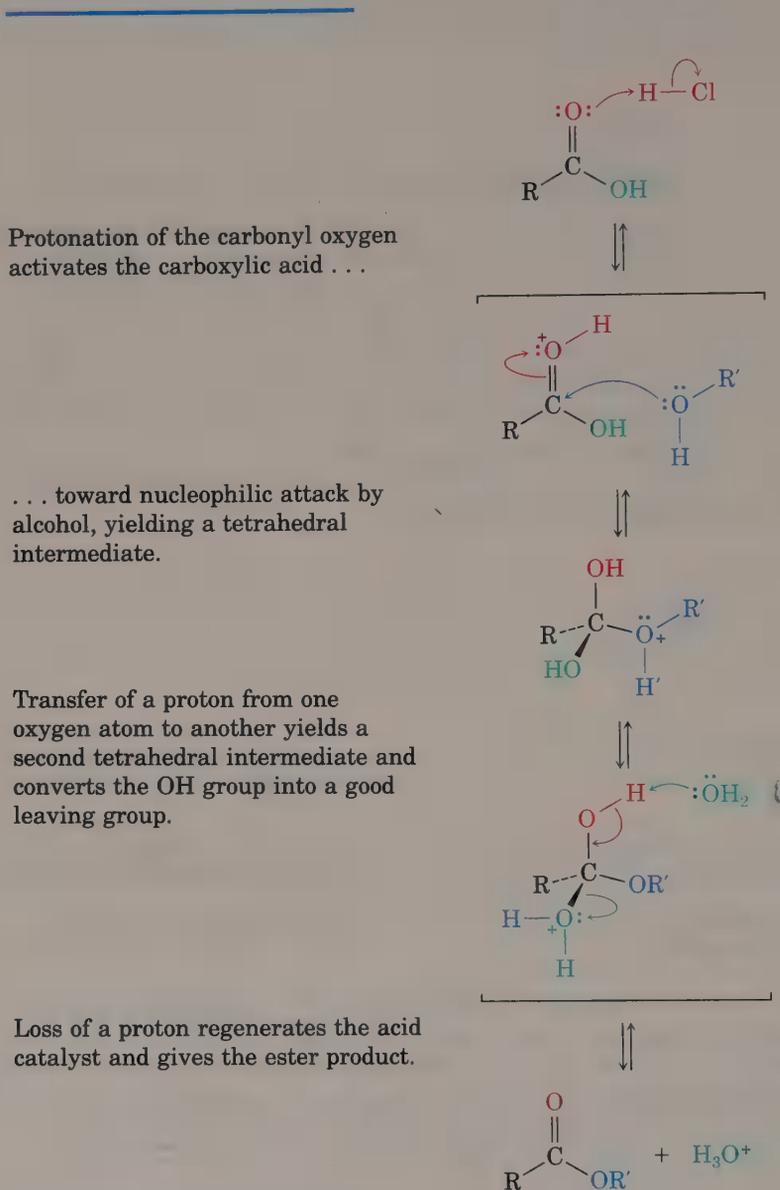


The Fischer esterification reaction, whose mechanism is shown in Figure 21.6, is a nucleophilic acyl substitution reaction carried out under acidic conditions. Although free carboxylic acids are not reactive enough to be attacked by most nucleophiles, they can be made much more reactive in the presence of a strong mineral acid such as HCl or  $\text{H}_2\text{SO}_4$ . The mineral acid acts by protonating the carbonyl-group oxygen atom, thereby giving the carboxylic acid a positive charge and rendering it much more reactive toward nucleophilic attack by alcohol. Subsequent loss of water yields the ester product.

The net effect of Fischer esterification is substitution of an  $-\text{OH}$  group by  $-\text{OR}'$ . All steps are reversible, and the reaction can be driven in either

<sup>1</sup>Emil Fischer (1852–1919); b. Euskirchen, Germany; Ph.D. Strasbourg (Baeyer); professor, universities of Erlangen, Würzburg, and Berlin; Nobel Prize (1902).

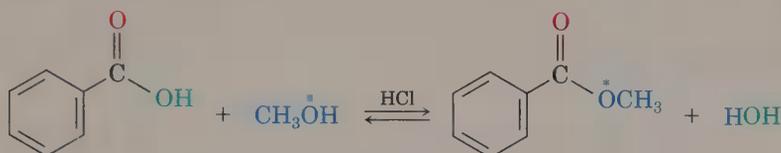
direction by choice of reaction conditions. Ester formation is favored when a large excess of alcohol is used as solvent, but carboxylic acid formation is favored when a large excess of water is present.



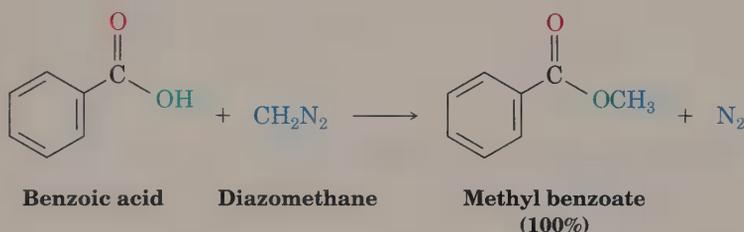
**Figure 21.6** Mechanism of Fischer esterification. The reaction is an acid-catalyzed nucleophilic acyl substitution process.

One of the best pieces of evidence in support of the mechanism shown in Figure 21.6 comes from isotopic-labeling experiments. When <sup>18</sup>O-labeled methanol reacts with benzoic acid under Fischer esterification conditions, the methyl benzoate produced is found to be <sup>18</sup>O labeled, but the water

produced is unlabeled. This experiment shows unequivocally that it is the CO–OH bond of the carboxylic acid that is cleaved, rather than the COO–H bond, and that it is the RO–H bond of the alcohol that is cleaved, rather than the R–OH bond.



A final method of ester synthesis is the reaction between a carboxylic acid and diazomethane,  $\text{CH}_2\text{N}_2$ , a reaction first described by von Pechmann<sup>2</sup> in 1894. The reaction takes place instantly at room temperature to give a high yield of the methyl ester. Though quite useful, this process does not involve a nucleophilic acyl substitution reaction, since it is the COO–H bond of the carboxylic acid that is broken rather than the CO–OH bond.



The diazomethane method of ester synthesis is ideal, since it occurs under mild, neutral conditions and gives nitrogen gas as the only by-product. Unfortunately, diazomethane is both toxic and explosive and should be handled only in small amounts by skilled persons.

PROBLEM.....

21.5 Show how you would prepare these esters:

(a) Butyl acetate

(b) Methyl butanoate

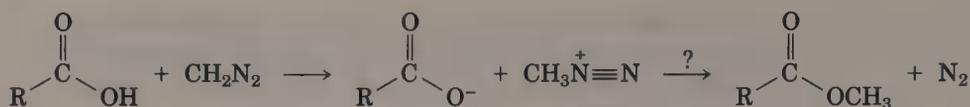
PROBLEM.....

21.6 If 5-hydroxypentanoic acid is treated with acid catalyst, an intramolecular esterification reaction occurs. What is the structure of the product? (*Intramolecular* means within the same molecule.)

PROBLEM.....

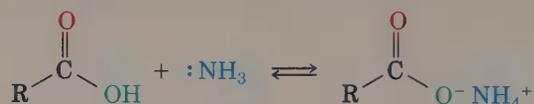
21.7 The first step in the reaction of a carboxylic acid with diazomethane is a proton transfer to give the methyldiazonium cation ( $\text{CH}_3\text{N}_2^+$ ) and carboxylate anion. What kind of mechanism accounts for the second step?

<sup>2</sup>Hans von Pechmann (1850–1902); b. Nuremberg; Ph.D. Greiswald; professor, universities of Munich and Tübingen.



## CONVERSION INTO AMIDES

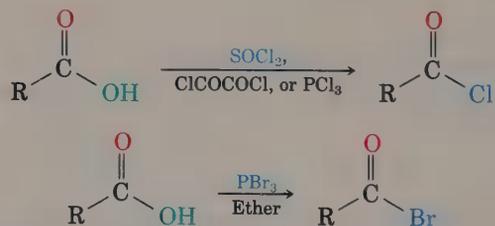
Amides are carboxylic acid derivatives in which the acid hydroxyl group has been replaced by a nitrogen substituent,  $-\text{NH}_2$ ,  $-\text{NHR}$ , or  $-\text{NR}_2$ . Amides are difficult to prepare directly from reaction between amines and carboxylic acids, because amines are bases that convert acidic carboxyl groups into their carboxylate anions. Since the carboxylate anion has a negative charge, it is no longer electrophilic and no longer likely to be attacked by nucleophiles.



## 21.5 Chemistry of Acid Halides

### PREPARATION OF ACID HALIDES

Acid chlorides are prepared from carboxylic acids by reaction with thionyl chloride ( $\text{SOCl}_2$ ), oxalyl chloride ( $\text{ClCOCOCl}$ ), or phosphorus trichloride ( $\text{PCl}_3$ ), as we saw in the previous section. Reaction of an acid with phosphorus tribromide ( $\text{PBr}_3$ ) yields the acid bromide.

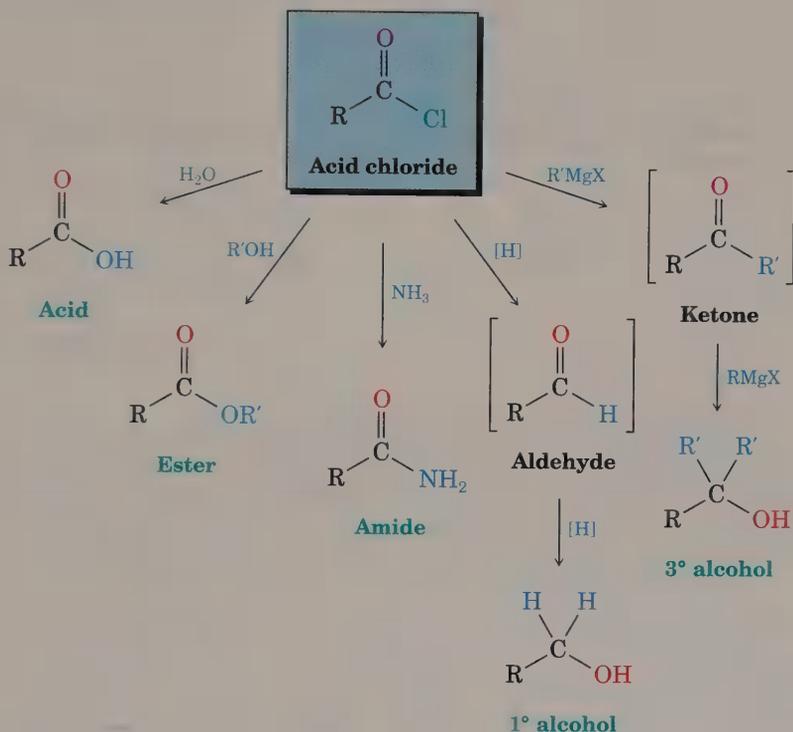


### REACTIONS OF ACID HALIDES

Acid halides are among the most reactive of carboxylic acid derivatives and can be converted into a variety of other kinds of compounds. For example, we've already seen the value of acid chlorides in preparing aryl alkyl ketones via the Friedel-Crafts reaction (Section 16.4).

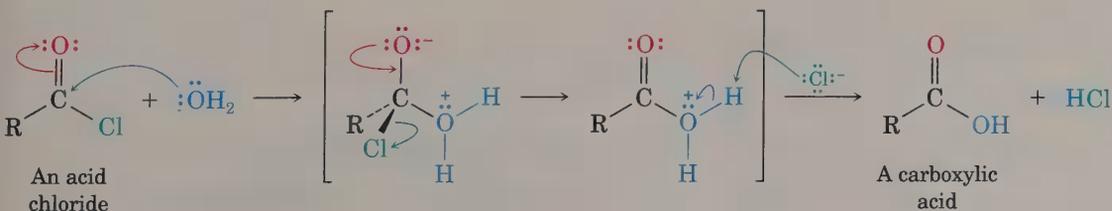


Most acid halide reactions occur by a nucleophilic acyl substitution mechanism. As illustrated in Figure 21.7, the halogen can be replaced by  $-\text{OH}$  to yield an acid, by  $-\text{OR}$  to yield an ester, or by  $-\text{NH}_2$  to yield an amide. In addition, the reduction of acid halides yields primary alcohols, and reaction with Grignard reagents yields tertiary alcohols. Although the reactions in Figure 21.7 are illustrated only for acid chlorides, they also take place with other acid halides.



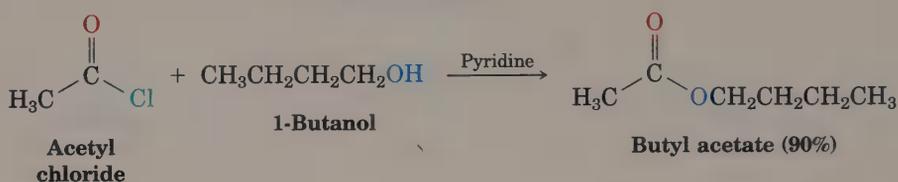
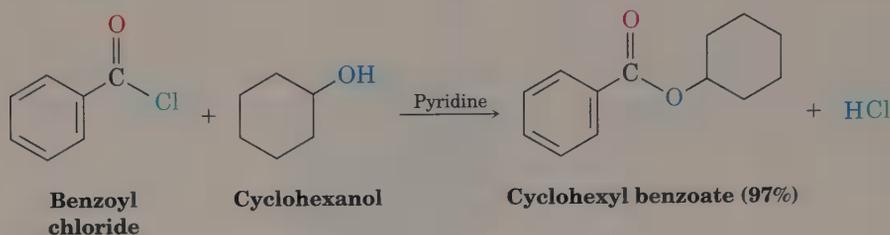
**Figure 21.7** Some nucleophilic acyl substitution reactions of acid chlorides.

**Hydrolysis: Conversion of Acid Halides into Acids** Acid chlorides react with water to yield carboxylic acids. This hydrolysis reaction is a typical nucleophilic acyl substitution process initiated by attack of water on the acid chloride carbonyl group. The initially formed tetrahedral intermediate undergoes elimination of chloride ion and loss of a proton to give the product carboxylic acid plus  $\text{HCl}$ .



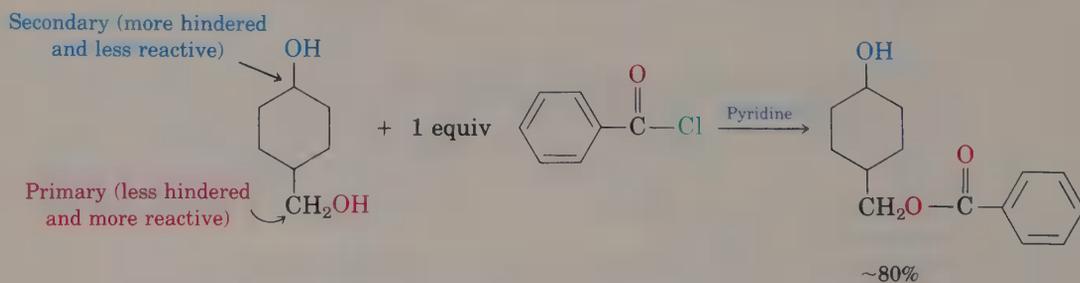
Since HCl is generated during the hydrolysis, the reaction is often carried out in the presence of pyridine or NaOH as a base to scavenge the HCl and prevent it from causing side reactions.

**Alcoholysis: Conversion of Acid Halides into Esters** Acid chlorides react with alcohols to yield esters in a process analogous to their reaction with water to yield acids.



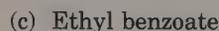
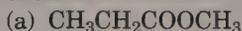
As with hydrolysis, alcoholysis reactions are usually carried out in the presence of pyridine or NaOH to react with the HCl formed and prevent it from causing side reactions. If this were not done, the HCl might react with the alcohol to form an alkyl chloride, or it might add to a carbon-carbon double bond if one were present elsewhere in the molecule.

The esterification of alcohols with acid chlorides is strongly affected by steric hindrance. Bulky groups on either reaction partner slow down the rate of reaction considerably, resulting in a reactivity order among alcohols of primary > secondary > tertiary. As a result, it's often possible to esterify an unhindered alcohol selectively in the presence of a more hindered one. This can be important in complex synthesis where it is often necessary to distinguish chemically between similar functional groups. For example:



**PROBLEM** .....

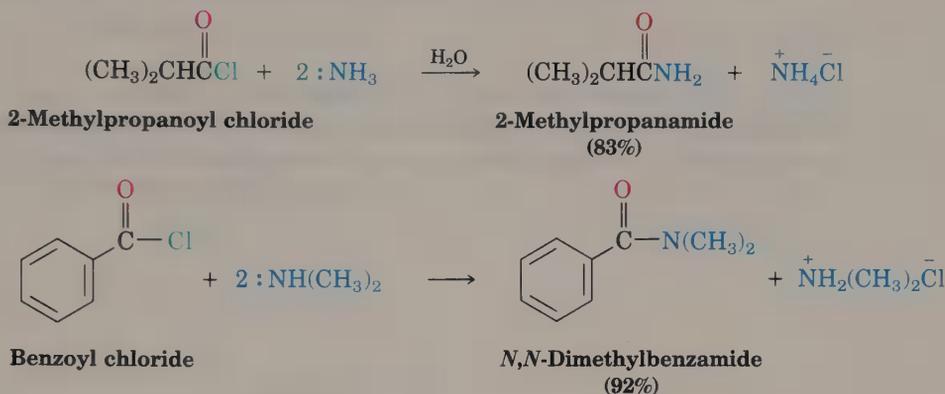
**21.8** How might you prepare these esters using a nucleophilic acyl substitution reaction of an acid chloride?



## PROBLEM.....

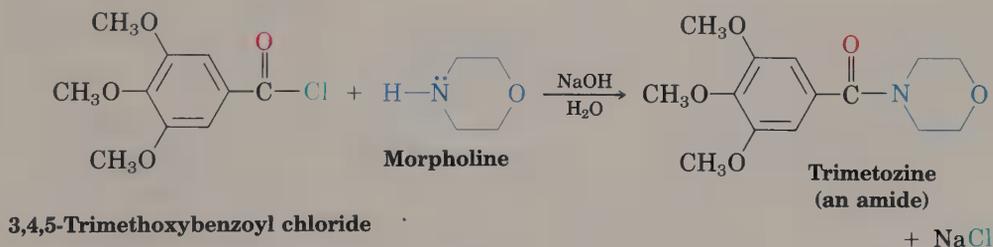
- 21.9 Which method would you choose if you wanted to prepare cyclohexyl benzoate: Fischer esterification or reaction of an acid chloride with an alcohol? Explain.
- .....

**Aminolysis: Conversion of Acid Halides into Amides** Acid chlorides react rapidly with ammonia and with amines to give amides. The reaction is rapid, and yields are usually excellent. Both mono- and disubstituted amines can be used, but not trisubstituted amines.



Since HCl is formed during the reaction, two equivalents of the amine must be used. One equivalent reacts with the acid chloride, and one equivalent reacts with HCl to form an ammonium chloride salt. If, however, the amine component is valuable, amide synthesis is often carried out using one equivalent of the desired amine plus one equivalent of an inexpensive base such as NaOH.

Aminolysis reactions carried out with NaOH present are sometimes referred to as Schotten–Baumann reactions after their discoverers.<sup>3</sup> For example, the medically useful sedative trimetozine is prepared by reaction of 3,4,5-trimethoxybenzoyl chloride with the amine morpholine in the presence of one equivalent of NaOH.



## PROBLEM.....

- 21.10 Write the steps in the mechanism of the reaction between 3,4,5-trimethoxybenzoyl chloride and morpholine to form trimetozine.

<sup>3</sup>Carl Schotten (1853–1910); b. Marburg, Germany; Ph.D. Berlin (Hofmann); professor, University of Berlin.

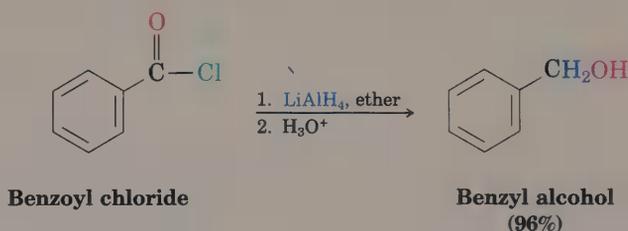
## PROBLEM .....

- 21.11 Trisubstituted amines such as triethylamine can be used in place of NaOH to scavenge HCl during aminolysis reactions. Why doesn't triethylamine react with acid chlorides to yield amides?

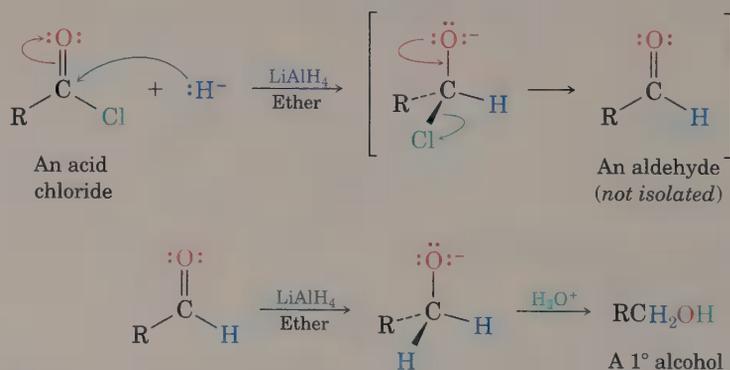
## PROBLEM .....

- 21.12 How could you prepare these amides using an acid chloride and an amine or ammonia?  
 (a)  $\text{CH}_3\text{CH}_2\text{CONHCH}_3$  (b) *N,N*-Diethylbenzamide (c) Propanamide

**Reduction: Conversion of Acid Chlorides into Alcohols** Acid chlorides are reduced by lithium aluminum hydride to yield primary alcohols. The reaction is of little practical value, however, since the parent carboxylic acids are generally more readily available and are themselves reduced by  $\text{LiAlH}_4$  to yield alcohols.

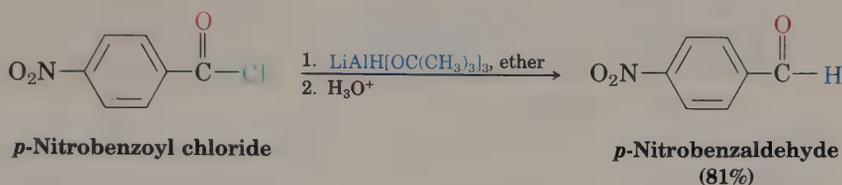
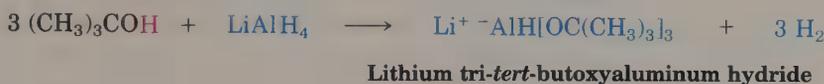


Reduction occurs via a typical nucleophilic acyl substitution mechanism in which a hydride ion ( $\text{H}^-$ ) attacks the carbonyl group, yielding a tetrahedral intermediate that expels chloride ion. The net effect is a substitution of  $-\text{H}$  for  $-\text{Cl}$ , yielding an aldehyde, which is then rapidly reduced by  $\text{LiAlH}_4$  in a second step to yield the primary alcohol.

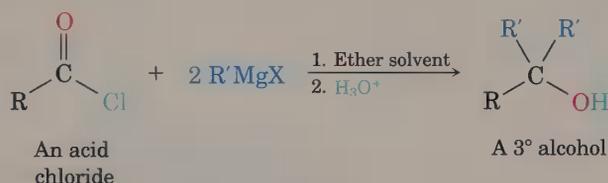


The aldehyde intermediate can be isolated if the less powerful hydride reducing agent, lithium tri-*tert*-butoxyaluminum hydride, is used. This reagent, which is obtained from reaction of  $\text{LiAlH}_4$  with three equivalents of

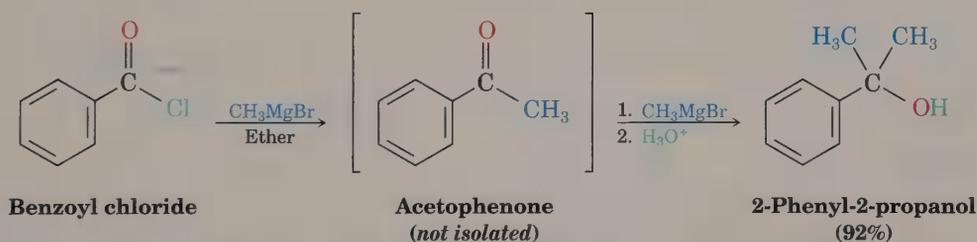
*tert*-butyl alcohol, is particularly effective for carrying out the partial reduction of acid chlorides to aldehydes:



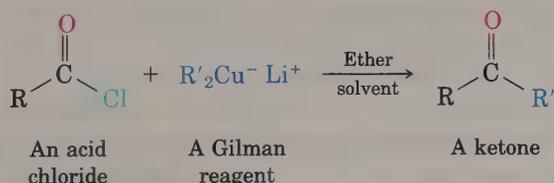
**Reaction of Acid Chlorides with Organometallic Reagents** Grignard reagents react with acid chlorides to yield tertiary alcohols in which two of the substituents are identical:



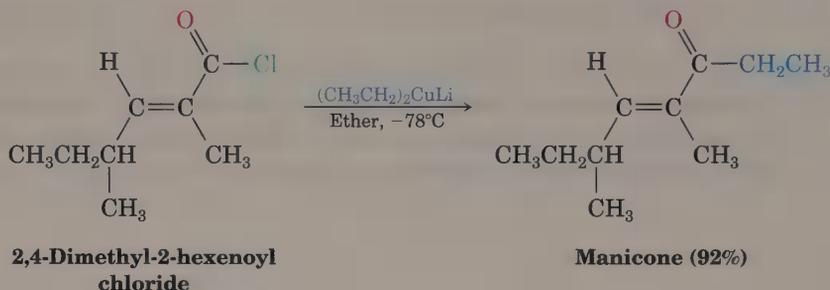
The mechanism for this Grignard reaction is similar to that for  $\text{LiAlH}_4$  reduction. The first equivalent of Grignard reagent attacks the acid chloride. Loss of chloride ion from the tetrahedral intermediate then yields a ketone intermediate, which reacts with a second equivalent of the organometallic reagent to produce an alcohol:



The ketone intermediate can't usually be isolated in Grignard reactions because addition of the second equivalent of organomagnesium reagent occurs so rapidly. Ketones can, however, be isolated from the reaction of acid chlorides with diorganocopper (Gilman) reagents (Section 19.4):



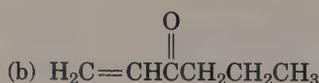
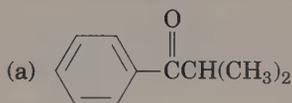
Reactions between acid chlorides and a wide variety of lithium diorganocoppers, including dialkyl-, diaryl-, and dialkenylcoppers, are possible. Despite their apparent similarity to Grignard reactions, however, these diorganocopper reactions are almost certainly not typical nucleophilic acyl substitution processes. Rather, it's thought that diorganocopper reactions occur via a radical pathway. The reactions are generally carried out at  $-78^{\circ}\text{C}$  in ether solution, and yields are often excellent. For example, manicone, a substance secreted by male ants to coordinate ant pairing and mating, has been synthesized by reaction of lithium diethylcopper with (*E*)-2,4-dimethyl-2-hexenoyl chloride:



Note that lithium diorganocopper reagents react only with acid chlorides. Acids, esters, anhydrides, and amides are inert to diorganocopper reagents.

PROBLEM .....

- 21.13 Show how these ketones might be prepared by reaction of an acid chloride with a lithium diorganocopper reagent.



PROBLEM .....

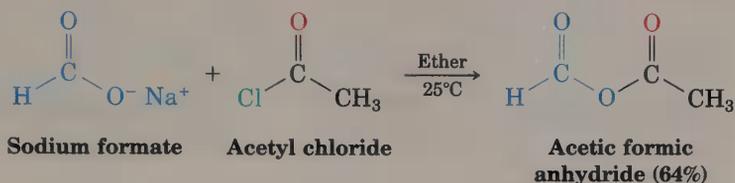
- 21.14 Ketones can be prepared by reaction of acid chlorides with one equivalent of Grignard reagent if the Grignard reagent is slowly added at low temperature to a solution of the acid chloride, rather than vice versa. How can you explain this observation?
- .....

## 21.6 Chemistry of Acid Anhydrides

### PREPARATION OF ACID ANHYDRIDES

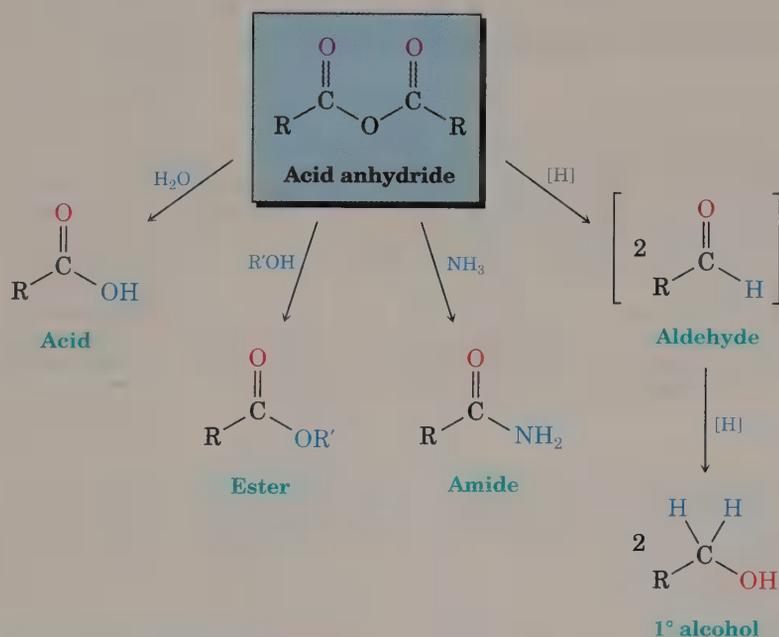
The most general method for preparing acid anhydrides is by nucleophilic acyl substitution reaction of an acid chloride with a carboxylate anion. Both

symmetrical and unsymmetrical acid anhydrides can be prepared in this way.



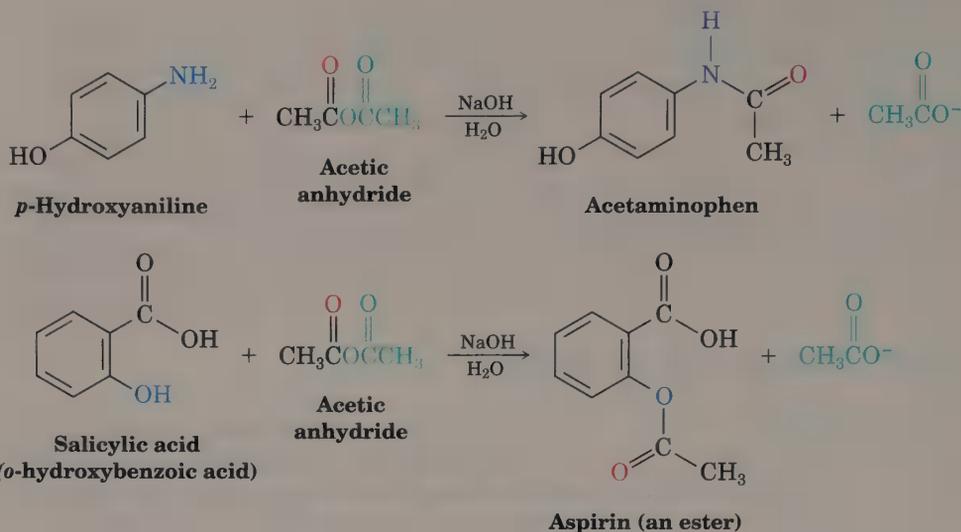
## REACTIONS OF ACID ANHYDRIDES

The chemistry of acid anhydrides is similar to that of acid chlorides. Although anhydrides react more slowly than acid chlorides, the kinds of reactions the two groups undergo are the same. Thus, acid anhydrides react with water to form acids, with alcohols to form esters, with amines to form amides, and with  $\text{LiAlH}_4$  to form primary alcohols (Figure 21.8).



**Figure 21.8** Some reactions of acid anhydrides.

Acetic anhydride is often used to prepare acetate esters of complex alcohols and to prepare substituted acetamides from amines. For example, acetaminophen, a drug used in headache remedies, can be prepared by reaction of *p*-hydroxyaniline with acetic anhydride. Aspirin (acetylsalicylic acid) is prepared similarly by the acetylation of *o*-hydroxybenzoic acid (salicylic acid) with acetic anhydride.



Notice in these two examples that only “half” of the anhydride molecule is used; the other half acts as the leaving group during the nucleophilic acyl substitution step and produces carboxylate anion as a by-product. Thus, anhydrides are inefficient to use, and acid chlorides are normally preferred for introducing acyl substituents other than acetyl groups.

PROBLEM .....

- 21.15 What product would you expect to obtain from reaction of one equivalent of methanol with a cyclic anhydride such as phthalic anhydride? What is the fate of the second “half” of the anhydride in such cases?

PROBLEM .....

- 21.16 Write the steps involved in the mechanism of reaction between *p*-hydroxyaniline and acetic anhydride to prepare acetaminophen.

PROBLEM .....

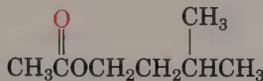
- 21.17 Why is one equivalent of a base such as sodium hydroxide required for the reaction between an amine and an anhydride to go to completion? What would happen if no base were present?
- .....

## 21.7 Chemistry of Esters

Esters are among the most important and widespread of all naturally occurring compounds. Many simple esters are pleasant-smelling liquids that are responsible for the fragrant odors of fruits and flowers. For example, methyl butanoate has been isolated from pineapple oil, and isopentyl acetate is a constituent of banana oil. The ester linkage is also present in animal fats and many biologically important molecules.



Methyl butanoate  
(from pineapples)

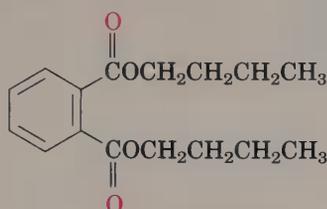


Isopentyl acetate  
(from bananas)



A fat  
(R = C<sub>12-18</sub> chains)

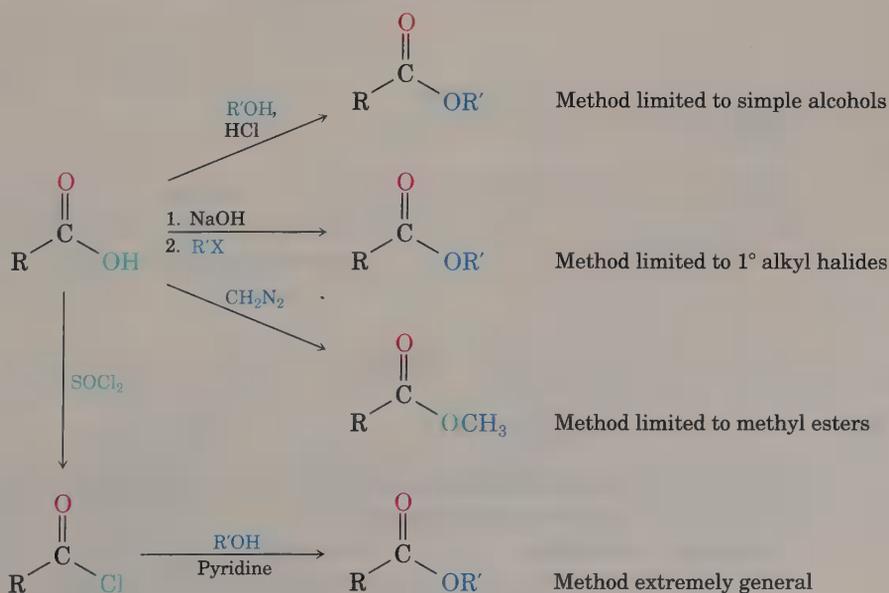
The chemical industry uses esters for a variety of purposes. For example, ethyl acetate is a common solvent found in nail-polish remover, and dialkyl phthalates are used as plasticizers to keep plastics from turning brittle.



Dibutyl phthalate (a plasticizer)

## PREPARATION OF ESTERS

Esters are usually prepared from acids by the methods already discussed. Thus, carboxylic acids are converted directly into esters by S<sub>N</sub>2 reaction of a carboxylate salt with a primary alkyl halide, by Fischer esterification of a carboxylic acid with a simple alcohol in the presence of a mineral acid catalyst, or by reaction of a carboxylic acid with diazomethane (Section 21.4). In addition, acid chlorides are converted into esters by treatment with an alcohol in the presence of base (Section 21.5).



## REACTIONS OF ESTERS

Esters show the same kinds of chemistry that we've seen for other acid derivatives, but they are less reactive toward nucleophiles than either acid chlorides or anhydrides are. Figure 21.9 shows some general reactions of esters. All of these reactions are equally applicable to both acyclic and cyclic esters (**lactones**).

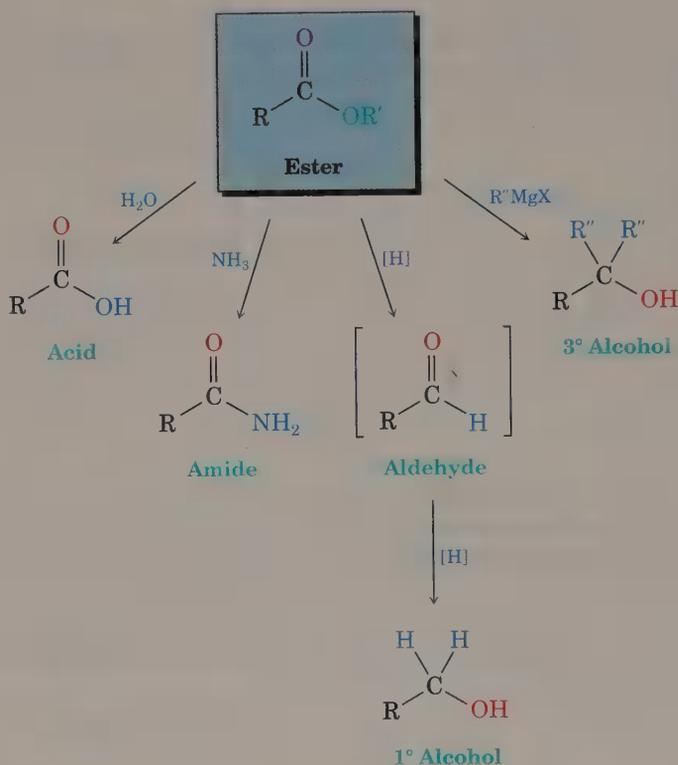
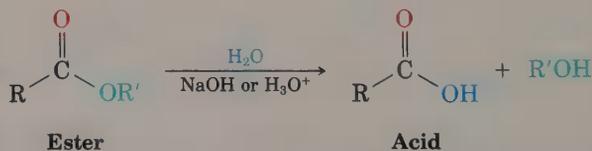


Figure 21.9 Some reactions of esters.

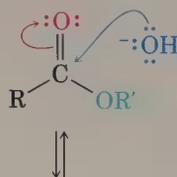
**Hydrolysis: Conversion of Esters into Carboxylic Acids** Esters are hydrolyzed either by aqueous base or by aqueous acid to yield carboxylic acid plus alcohol:



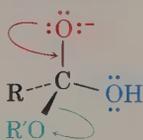
Ester hydrolysis in basic solution is called **saponification**, after the Latin *sapo*, “soap.” As we’ll see in Section 28.2, the boiling of animal fat with extract of wood ash to make soap is indeed a saponification, since wood ash contains alkali and fats have ester linkages.

Ester hydrolysis occurs through the typical nucleophilic acyl substitution pathway shown in Figure 21.10, in which hydroxide ion nucleophile adds to the ester carbonyl group to give a tetrahedral intermediate. Loss of alkoxide ion then gives a carboxylic acid, which is deprotonated to give the carboxylate salt. The free acid is then obtained by addition of aqueous HCl to protonate the carboxylate.

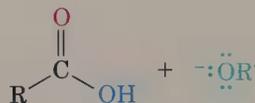
Nucleophilic addition of hydroxide ion to the ester carbonyl group gives the usual tetrahedral alkoxide intermediate.



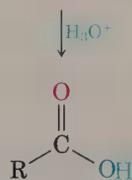
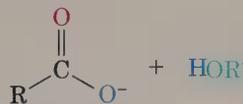
Elimination of alkoxide ion then generates the carboxylic acid.



Alkoxide ion abstracts the acidic proton from the carboxylic acid and yields a carboxylate ion.

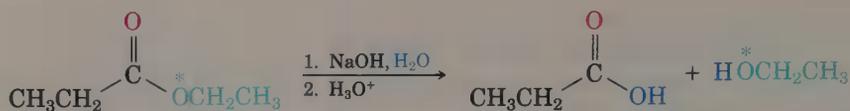


Protonation of the carboxylate ion by addition of aqueous mineral acid gives the free carboxylic acid.



**Figure 21.10** Mechanism of base-induced ester hydrolysis.

One of the most elegant experiments in support of this mechanism involves isotope labeling. When ethyl propanoate labeled with  $^{18}\text{O}$  in the ether-type oxygen is hydrolyzed in aqueous sodium hydroxide, the  $^{18}\text{O}$  label shows up exclusively in the ethanol product. None of the label remains with the propanoic acid, indicating that saponification occurs by cleavage of the acyl-oxygen bond ( $\text{RCO}-\text{OR}'$ ) rather than the alkyl-oxygen bond ( $\text{RCOO}-\text{R}'$ ). This result is just what we would expect, based on our knowledge of the nucleophilic acyl substitution mechanism.



Acidic hydrolysis of esters can occur by more than one mechanism, depending on the structure of the substrate. The usual pathway, however, is just the reverse of the Fischer esterification reaction (Section 21.4). The ester is first activated toward nucleophilic attack by protonation of the carboxyl oxygen atom. Nucleophilic attack by water, followed by transfer of a proton and elimination of alcohol, then yields the carboxylic acid (Figure 21.11).

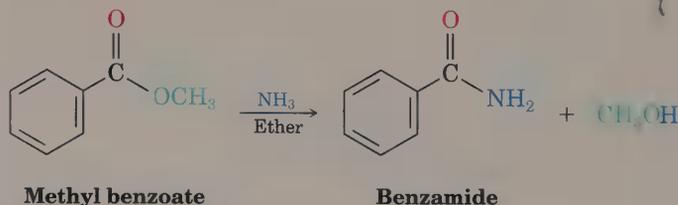
**PROBLEM** .....

- 21.18 How would you synthesize the  $^{18}\text{O}$ -labeled ethyl propanoate used in the mechanistic studies? Assume that  $^{18}\text{O}$ -labeled acetic acid is your only source of isotopic oxygen.

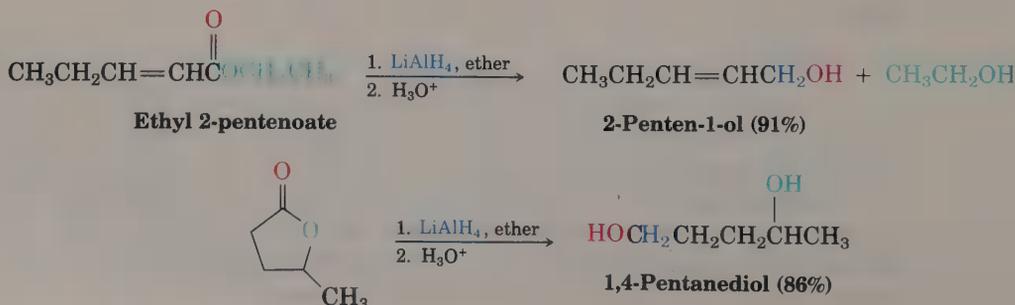
**PROBLEM** .....

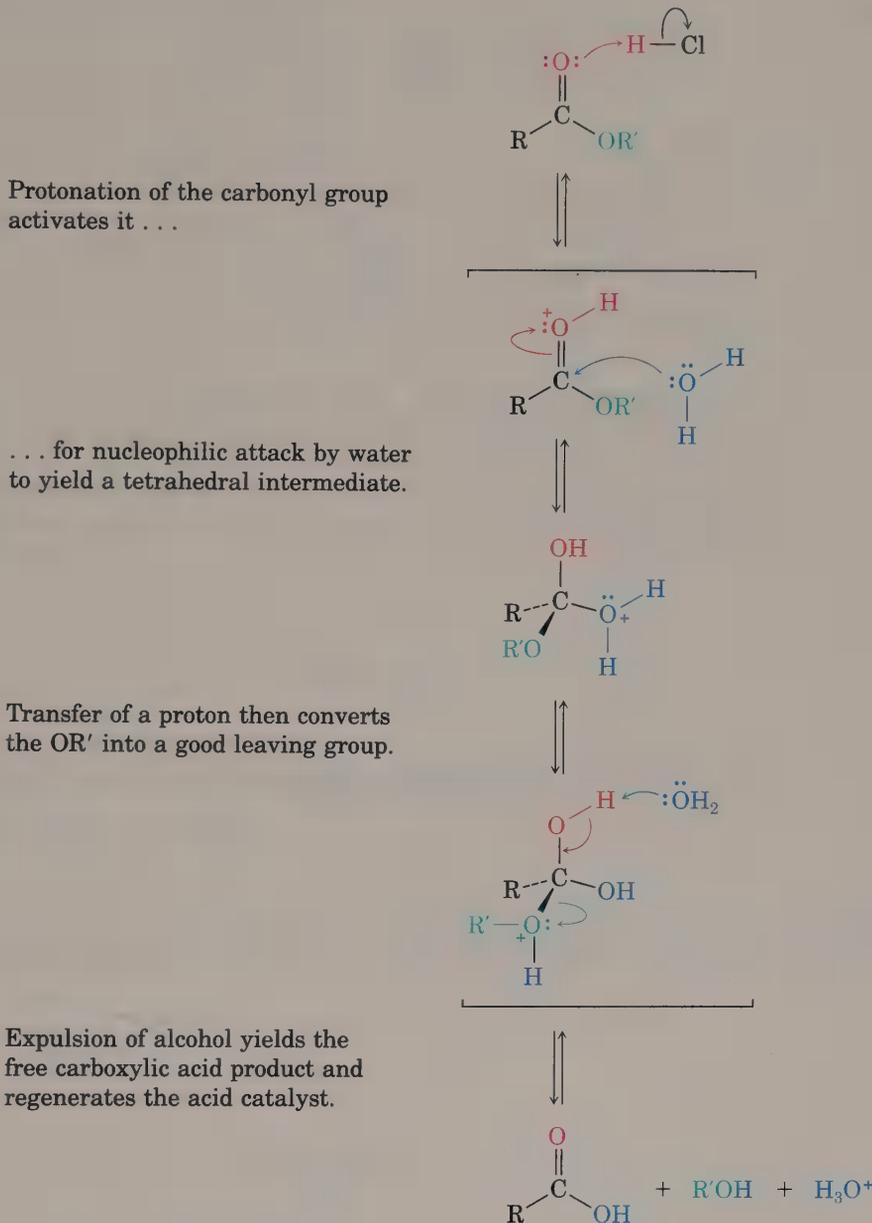
- 21.19 Why do you suppose saponification of esters is irreversible? In other words, why doesn't treatment of a carboxylic acid with alkoxide ion lead to ester formation?
- .....

**Aminolysis: Conversion of Esters into Amides** Esters react with ammonia and amines via a typical nucleophilic acyl substitution pathway to yield amides. The reaction is not often used, however, since higher yields are normally obtained by aminolysis of acid chlorides (Section 21.5).



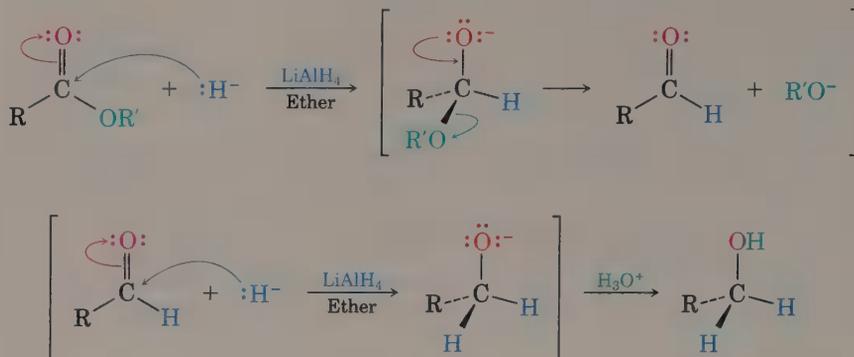
**Reduction: Conversion of Esters into Alcohols** Esters are easily reduced by treatment with lithium aluminum hydride to yield primary alcohols (Section 17.6).



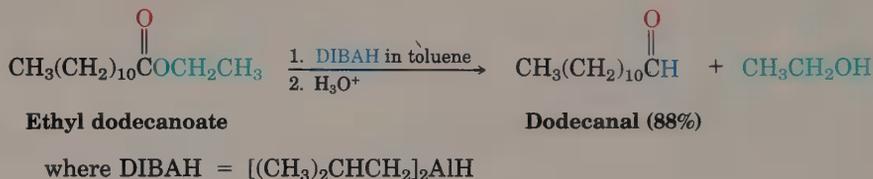


**Figure 21.11** Mechanism of acid-catalyzed ester hydrolysis. The forward reaction is a hydrolysis; the back-reaction is a Fischer esterification.

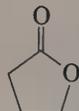
The mechanism of ester (and lactone) reductions is similar to that for acid chloride reduction. A hydride ion first adds to the carbonyl group, followed by elimination of alkoxide ion to yield an aldehyde intermediate. Further addition of hydride to the aldehyde gives the primary alcohol.



The aldehyde intermediate can be isolated if DIBAH (diisobutylaluminum hydride) is used as the reducing agent instead of  $\text{LiAlH}_4$ . Great care must be taken: Exactly one equivalent of hydride reagent must be used, and the reaction must be carried out at  $-78^\circ\text{C}$ . If these conditions are met, though, the DIBAH reduction of esters can be an excellent method of aldehyde synthesis.



**21.20** What product would you expect from the reaction of butyrolactone with DIBAH?

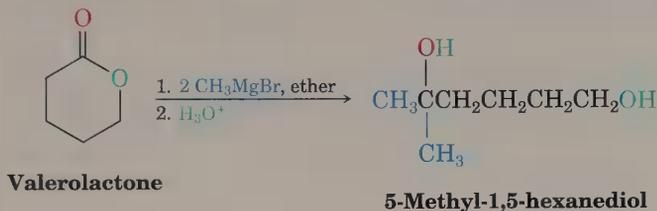
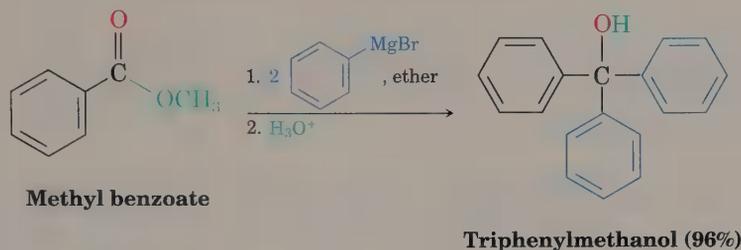


**Butyrolactone**

**PROBLEM**

**21.21** Show the products you would obtain by reduction of these esters with  $\text{LiAlH}_4$ .  
 (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)\text{COOCH}_3$       (b) Phenyl benzoate

**Reaction of Esters with Grignard Reagents** Esters and lactones react with two equivalents of Grignard reagent or organolithium reagent to yield tertiary alcohols (Section 17.7). The reactions occur readily and give excellent yields of products by the usual nucleophilic addition mechanism. For example:



**PROBLEM** .....

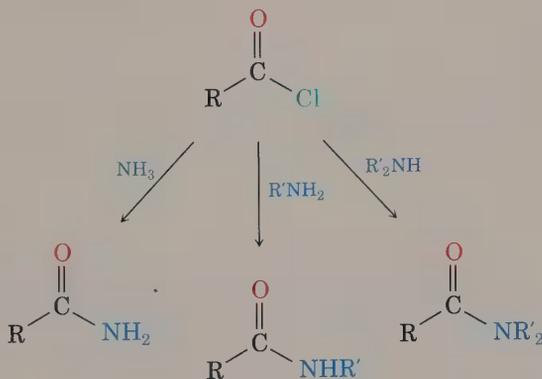
21.22 What ester and what Grignard reagent might you start with to prepare these alcohols?

- (a) 2-Phenyl-2-propanol      (b) 1,1-Diphenylethanol      (c) 3-Ethyl-3-heptanol
- .....

## 21.8 Chemistry of Amides

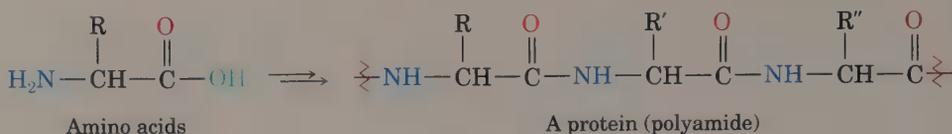
### PREPARATION OF AMIDES

Amides are usually prepared by reaction of an acid chloride with an amine. Ammonia, monosubstituted amines, and disubstituted amines all undergo this reaction (Section 21.5).



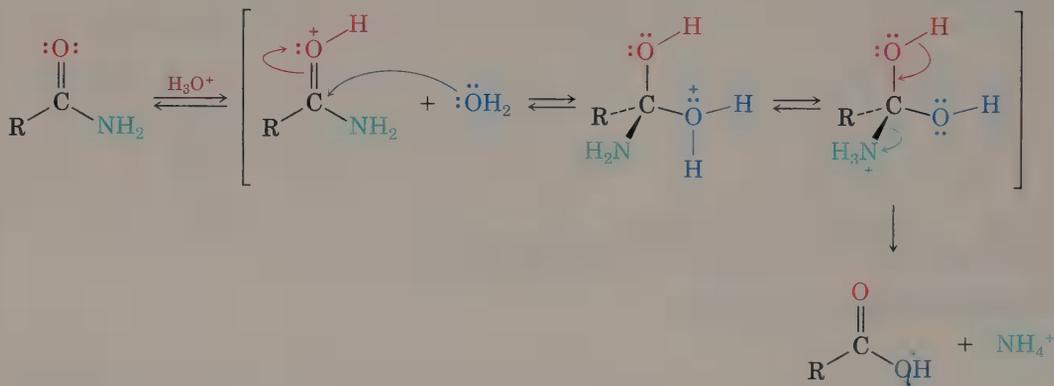
### REACTIONS OF AMIDES

Amides are much less reactive than acid chlorides, acid anhydrides, or esters. Thus, the amide linkage serves as the basic unit from which all proteins are made (Chapter 27).

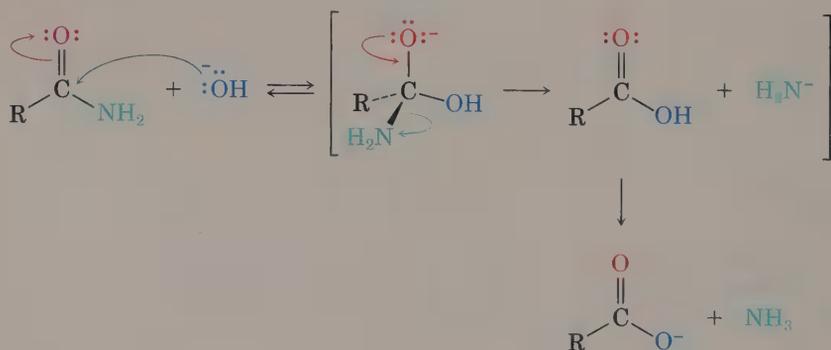


Amides undergo hydrolysis to yield carboxylic acids plus amine on heating in either aqueous acid or aqueous base. The conditions required for amide hydrolysis are more severe than those required for the hydrolysis of acid chlorides or esters, but the mechanisms are similar. The basic hydrolysis of amides yields an amine and a carboxylate ion as products and occurs by nucleophilic addition of hydroxide ion to the amide carbonyl group, followed by elimination of amide ion ( $^-:\text{NH}_2$ ). The acidic hydrolysis reaction occurs by nucleophilic addition of water to the protonated amide, followed by loss of ammonia.

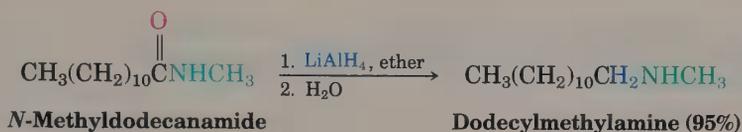
#### Acidic hydrolysis



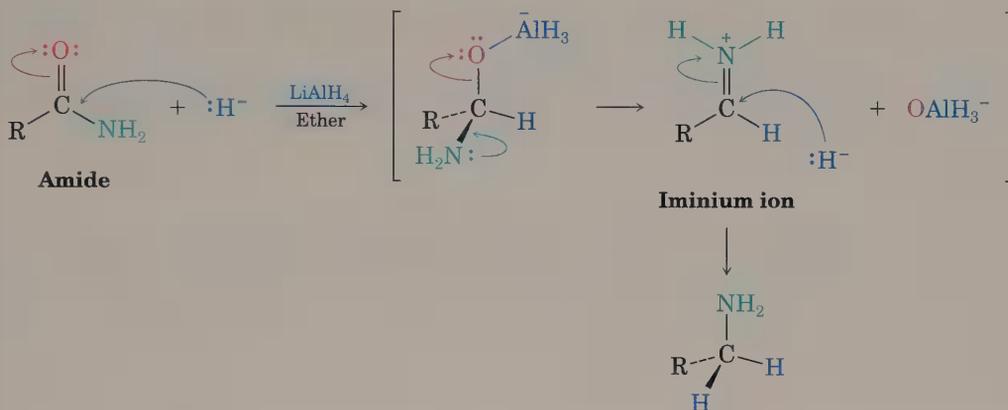
#### Basic hydrolysis



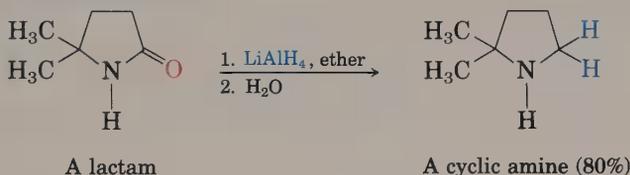
Like other carboxylic acid derivatives, amides can be reduced by  $\text{LiAlH}_4$ . The product of this reduction, however, is an amine rather than an alcohol. The net effect of amide reduction reaction is to convert the amide carbonyl group into a methylene group ( $\text{C}=\text{O} \rightarrow \text{CH}_2$ ). This kind of reaction is specific for amides and does not occur with other carboxylic acid derivatives.



Amide reduction occurs by initial nucleophilic addition of hydride ion to the amide carbonyl group, followed by expulsion of the oxygen atom as an aluminate anion to give an iminium ion intermediate. The intermediate iminium ion is then further reduced by  $\text{LiAlH}_4$  to yield the amine.



Lithium aluminum hydride reduction is equally effective with both acyclic and cyclic amides (**lactams**). Lactam reductions provide cyclic amines in good yield and constitute a valuable method of synthesis.

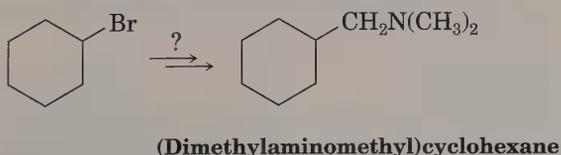


PROBLEM .....

- 21.23 How would you convert *N*-ethylbenzamide to these products?  
 (a) Benzoic acid    (b) Benzyl alcohol    (c)  $\text{C}_6\text{H}_5\text{CH}_2\text{NHCH}_2\text{CH}_3$

PROBLEM .....

- 21.24 How would you use the reaction between an amide and  $\text{LiAlH}_4$  as the key step in going from bromocyclohexane to (dimethylaminomethyl)cyclohexane? Formulate all steps involved in the reaction sequence.

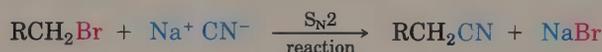


## 21.9 Chemistry of Nitriles

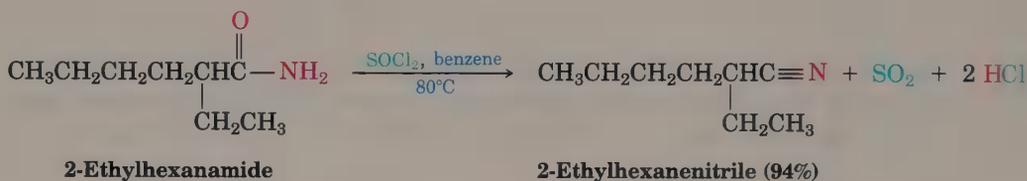
Nitriles are not related to carboxylic acids in the same sense that acyl derivatives are, yet the chemistry of nitriles and carboxylic acids is so entwined that the two classes of compounds should be considered together.

### PREPARATION OF NITRILES

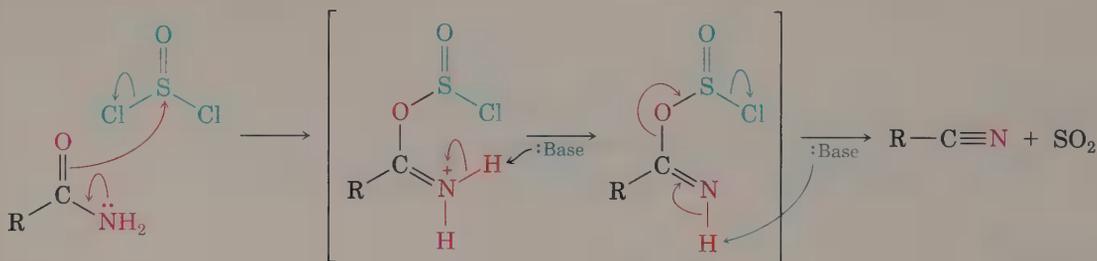
The simplest method of nitrile preparation is  $S_N2$  reaction of cyanide ion with a primary alkyl halide, a reaction discussed in Section 20.6. This method is limited by the usual  $S_N2$  steric constraints to the synthesis of  $\alpha$ -unsubstituted nitriles,  $RCH_2CN$ .



Another excellent method for preparing nitriles is by dehydration of a primary amide. Thionyl chloride is often used to effect this reaction, although other dehydrating agents such as  $P_2O_5$ ,  $POCl_3$ , and acetic anhydride can also be used.



Amide dehydration occurs by initial reaction on the amide oxygen atom, followed by an elimination reaction.

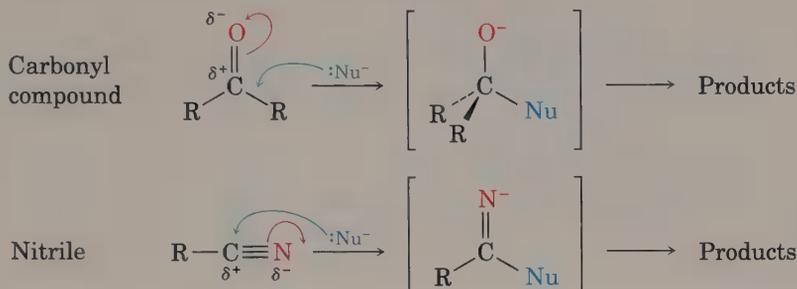


Although both methods of nitrile synthesis— $S_N2$  displacement by cyanide ion on an alkyl halide and amide dehydration—are useful, the synthesis from amides is more general since it is not limited by steric hindrance.

### REACTIONS OF NITRILES

The chemistry of nitriles is similar in many respects to the chemistry of carbonyl compounds. Like carbonyl groups, nitriles are strongly polarized,

making the carbon atom electrophilic. Thus, nitriles are attacked by nucleophiles to yield  $sp^2$ -hybridized intermediate imine anions in a reaction analogous to the formation of an  $sp^3$ -hybridized alkoxide ion intermediate by nucleophilic addition to a carbonyl group.



The two most important reactions of nitriles are hydrolysis and reduction. In addition, nitriles can be partially reduced and hydrolyzed to yield aldehydes, and can be treated with Grignard reagents to yield ketones (Figure 21.12).

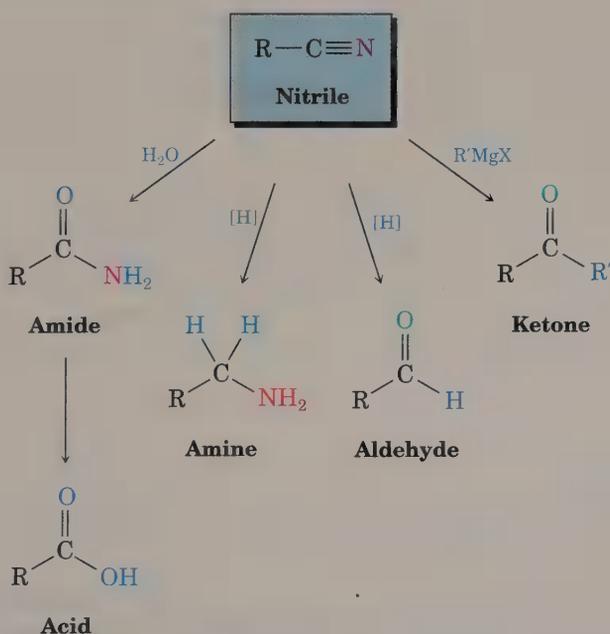
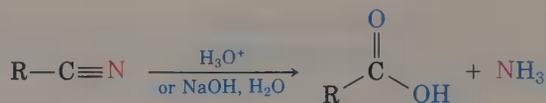
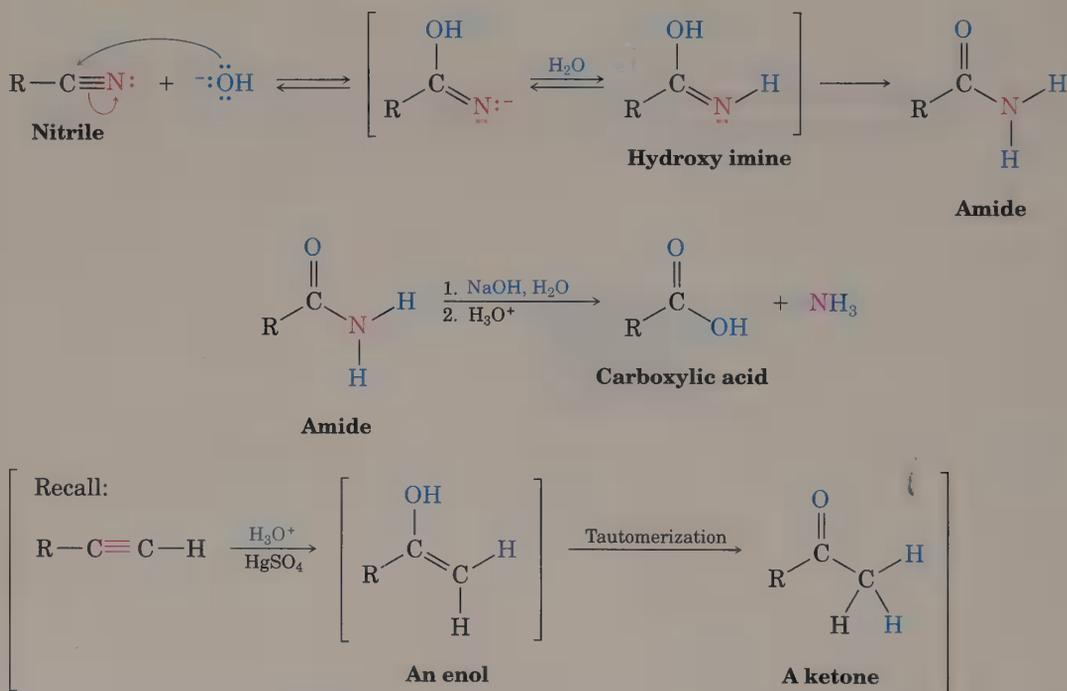


Figure 21.12 Some reactions of nitriles.

**Hydrolysis: Conversion of Nitriles into Acids** Nitriles are hydrolyzed in either acidic or basic aqueous solution to yield carboxylic acids and ammonia or an amine:



The mechanism of the alkaline hydrolysis involves nucleophilic addition of hydroxide ion to the polar  $\text{C}\equiv\text{N}$  bond in a manner analogous to that of nucleophilic addition to a polar carbonyl  $\text{C}=\text{O}$  bond. The initial product is a hydroxy imine, which is rapidly isomerized to an amide in a step similar to the isomerization of an enol to a ketone (Section 8.5). Further hydrolysis of the amide as discussed in the previous section then yields the carboxylic acid (Figure 21.13).



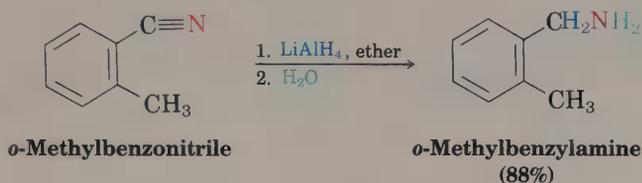
**Figure 21.13** Mechanism of amide hydrolysis to yield a carboxylic acid.

The conditions required for nitrile alkaline hydrolysis are severe ( $\text{KOH}$ ,  $200^\circ\text{C}$ ), and the amide intermediate can sometimes be isolated if milder conditions are used.

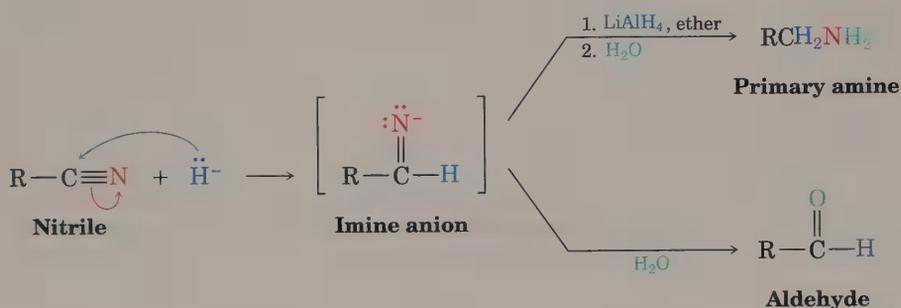
**PROBLEM** .....

- 21.25** Acid-catalyzed nitrile hydrolysis occurs by initial protonation of the nitrile nitrogen atom, followed by nucleophilic addition of water. Show all the steps involved in the acidic hydrolysis of a nitrile to yield a carboxylic acid.
- .....

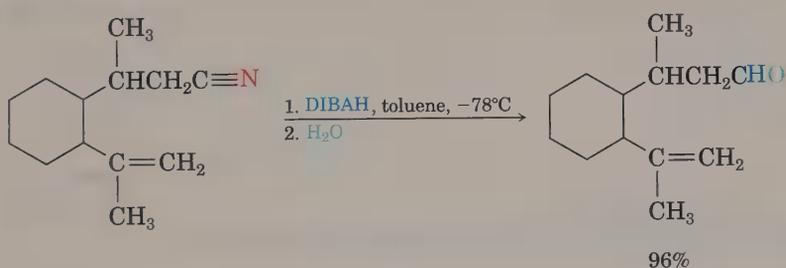
**Reduction: Conversion of Nitriles into Amines** Reaction of nitriles with lithium aluminum hydride gives primary amines in high yields. For example:



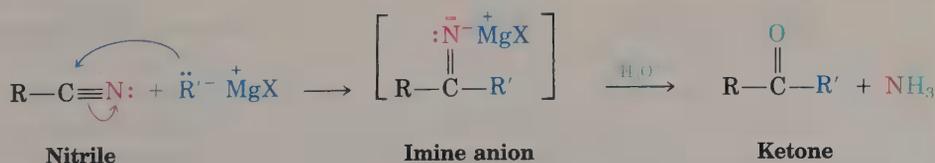
The reaction occurs by nucleophilic addition of hydride ion to the polar  $\text{C}\equiv\text{N}$  bond, yielding an imine anion, which undergoes further addition of a second equivalent of hydride to give the final product. If, however, a less powerful reducing agent such as DIBAH is used, the second addition of hydride does not occur, and the imine intermediate can be hydrolyzed to yield an aldehyde.



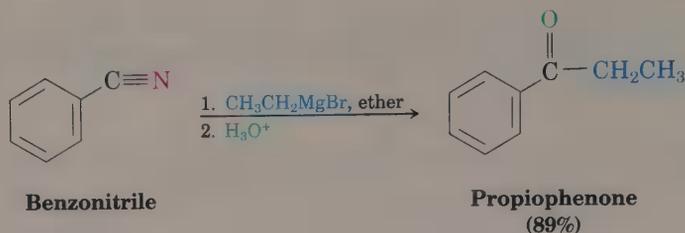
For example:



**Reaction of Nitriles with Organometallic Reagents** Grignard reagents add to nitriles, giving intermediate imine anions that can be hydrolyzed to yield ketones:

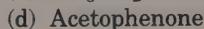
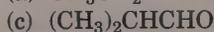
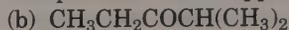
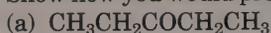


The reaction is similar to the DIBAH reduction of nitriles, except that the attacking nucleophile is a carbanion ( $\text{:R}'^-$ ) rather than a hydride ion. Yields are generally high. For example:

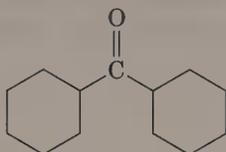


PROBLEM .....

21.26 Show how you would prepare these carbonyl compounds from an appropriate nitrile:



(e)



PROBLEM .....

21.27 How would you prepare 1-phenyl-2-butanone from benzyl bromide,  $\text{C}_6\text{H}_5\text{CH}_2\text{Br}$ ? More than one step is required.

PROBLEM .....

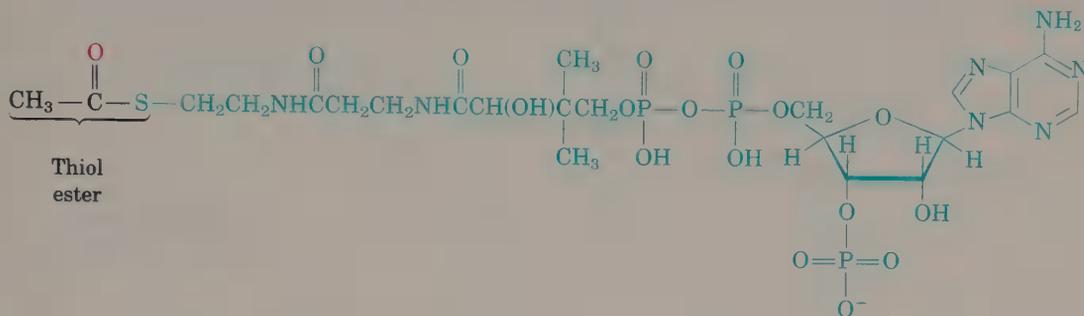
21.28 Why do you suppose only one equivalent of Grignard reagent adds to nitriles? Why doesn't a second Grignard reagent add to the initially formed imine anion?

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## 21.10 Thiol Esters: Biological Carboxylic Acid Derivatives

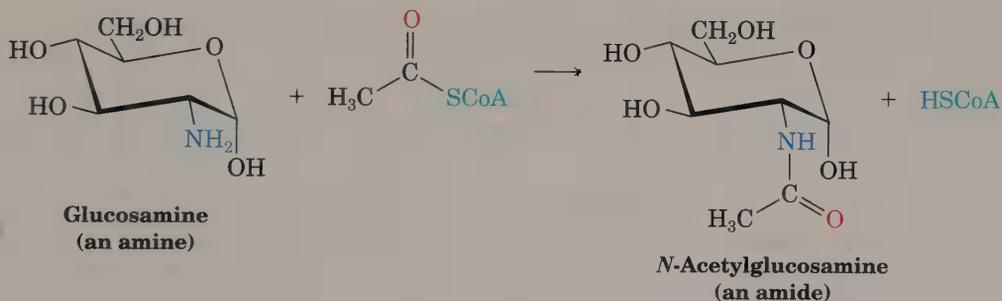
Nucleophilic acyl substitution reactions take place in living organisms just as they take place in the chemical laboratory. The same principles apply in both cases. Nature, however, uses **thiol esters**,  $\text{RCOSR}'$ —rather than acid chlorides or acid anhydrides—as reactive acyl derivatives. The  $\text{p}K_{\text{a}}$  of a typical alkane thiol ( $\text{RSH}$ ) is about 10, placing thiols midway between carboxylic acids ( $\text{p}K_{\text{a}} \approx 5$ ) and alcohols ( $\text{p}K_{\text{a}} \approx 16$ ) in acid strength. As a result, thiol esters are intermediate in reactivity between acid anhydrides and esters. They aren't so reactive that they hydrolyze rapidly, as anhydrides do, yet they're more reactive than normal esters toward nucleophilic attack.

Acetyl coenzyme A (usually abbreviated acetyl CoA; Figure 21.14) is the most common thiol ester found in nature. Acetyl CoA is a much more complex molecule than acetyl chloride or acetic anhydride, yet it serves exactly the same purpose as these simpler reagents do. Nature uses acetyl CoA as a reactive acylating agent in nucleophilic acyl substitution reactions.



**Figure 21.14** The structure of acetyl coenzyme A, abbreviated acetyl CoA or  $\text{CH}_3\text{COSCoA}$ .

As an example, *N*-acetylglucosamine, an important constituent of cell-surface membranes in mammals, is synthesized in nature by an aminolysis reaction between glucosamine and acetyl CoA:



## 21.11 Spectroscopy of Carboxylic Acid Derivatives

### INFRARED SPECTROSCOPY

All carbonyl-containing compounds have intense infrared absorptions in the range  $1650\text{--}1850\text{ cm}^{-1}$ . As shown in Table 21.2, the exact position of the absorption provides information about the nature of the specific kind of carbonyl group. For comparison, the absorptions of ketones, aldehydes, and acids are included in the table, along with values for carboxylic acid derivatives.



## PROBLEM.....

21.30 Propose structures for compounds having the following formulas and infrared absorptions:

- (a)  $C_3H_5N$ ,  $2250\text{ cm}^{-1}$  (b)  $C_6H_{12}O_2$ ,  $1735\text{ cm}^{-1}$   
 (c)  $C_4H_9NO$ ,  $1650\text{ cm}^{-1}$  (d)  $C_4H_5ClO$ ,  $1780\text{ cm}^{-1}$

## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Although  $^{13}C$  NMR is useful for determining the presence or absence of a carbonyl group in a molecule of unknown structure, precise information about the nature of the carbonyl group is difficult to obtain. Carbonyl carbon atoms show resonances in the range  $160\text{--}210\ \delta$ , as Table 21.3 shows.

TABLE 21.3 Positions of  $^{13}C$  NMR Absorptions in Some Carbonyl Compounds

Compound	Absorption ( $\delta$ )	Compound	Absorption ( $\delta$ )
Acetic acid	177.3	Acetic anhydride	166.9
Ethyl acetate	170.7	Acetonitrile	117.4
Acetyl chloride	170.3	Acetone	205.6
Acetamide	172.6	Acetaldehyde	201.0

Protons on the carbon next to a carbonyl group are slightly deshielded and absorb near  $2\ \delta$  in the  $^1H$  NMR spectrum. The exact nature of the carbonyl group can't be distinguished by  $^1H$  NMR, however, because all acyl derivatives absorb in the same range. Figure 21.15 shows the  $^1H$  NMR spectrum of ethyl acetate.

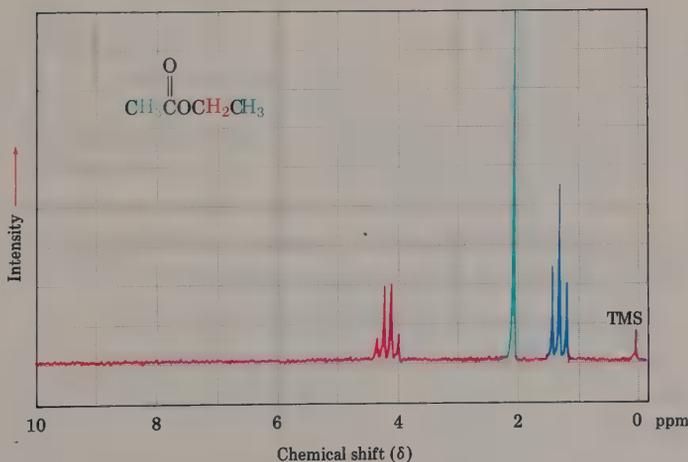
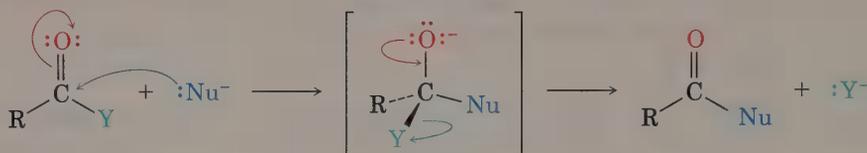


Figure 21.15 Proton NMR spectrum of ethyl acetate.

## 21.12 Summary and Key Words

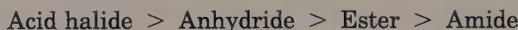
Carboxylic acids can be transformed into a variety of acid derivatives in which the acid  $-OH$  group has been replaced by other substituents. **Acid chlorides, acid anhydrides, esters, and amides** are the most important derivatives.

The chemistry of the different acid derivatives is dominated by a single general reaction type: the **nucleophilic acyl substitution reaction**. Mechanistically, these substitutions take place by addition of a nucleophile to the polar carbonyl group of the acid derivative, followed by expulsion of a leaving group from the tetrahedral intermediate.



where  $Y = Cl, Br, I$  (acid halide);  $OR$  (ester);  $OCOR$  (anhydride); or  $NH_2$  (amide)

The reactivity of an acid derivative toward substitution depends both on the steric environment near the carbonyl group and on the electronic nature of the substituent,  $Y$ . Thus, we find a reactivity order:



The most important reactions of carboxylic acid derivatives are substitution by water (**hydrolysis**) to yield an acid, by alcohols (**alcoholysis**) to yield an ester, by amines (**aminolysis**) to yield an amide, by hydride (**reduction**) to yield an alcohol, and by organometallic reagents (**Grignard reaction**) to yield an alcohol.

Nitriles undergo nucleophilic addition to the polar  $C\equiv N$  bond in the same way that carbonyl compounds do. The most important reactions of nitriles are their hydrolysis to carboxylic acids, their reduction to primary amines, their partial reduction to aldehydes, and their reaction with organometallic reagents to yield ketones.

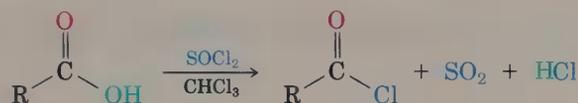
Nature employs nucleophilic acyl substitution reactions in the biosynthesis of many molecules, using **thiol esters** for the purpose. Acetyl coenzyme A (**acetyl CoA**) is a complex thiol ester that is employed in living systems to acetylate amines and alcohols.

Infrared spectroscopy is an extremely valuable tool for the structural analysis of acid derivatives. Acid chlorides, anhydrides, esters, amides, and nitriles all show characteristic infrared absorptions that can be used to identify these functional groups in unknowns. Carbon-13 NMR is useful for establishing the presence or absence of a carbonyl carbon atom but does not usually allow exact identification of the different functional groups.

## 21.13 Summary of Reactions

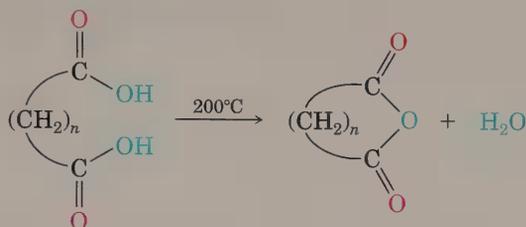
### 1. Reactions of carboxylic acids

#### a. Conversion into acid chlorides (Section 21.4)



The reaction can also be carried out using  $\text{PCl}_3$  or  $(\text{COCl})_2$ .

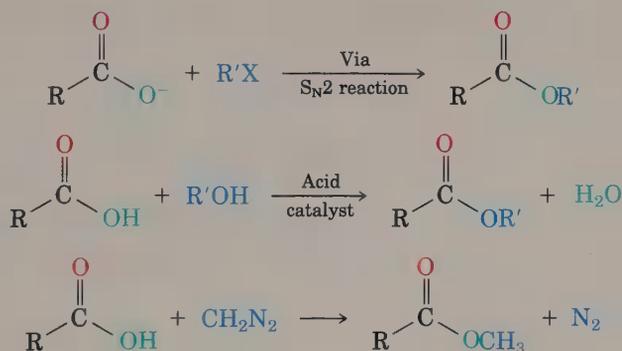
#### b. Conversion into cyclic acid anhydrides (Section 21.4)



where  $n = 2$  or  $3$

where  $n = 2$  or  $3$

#### c. Conversion into esters (Section 21.4)

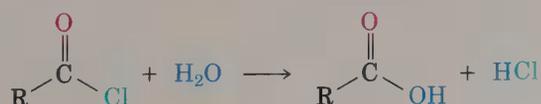


#### d. Conversion into amides (Section 21.4)

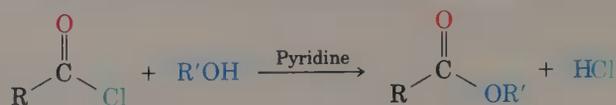


### 2. Reactions of acid chlorides

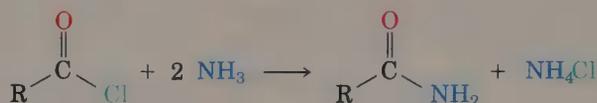
#### a. Hydrolysis to yield acids (Section 21.5)



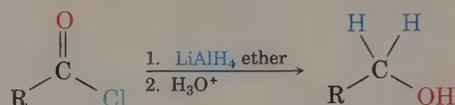
b. Alcoholysis to yield esters (Section 21.5)



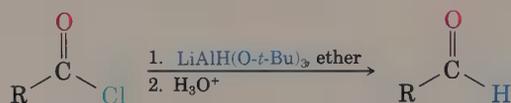
c. Aminolysis to yield amides (Section 21.5)



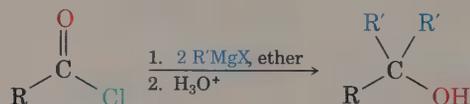
d. Reduction to yield primary alcohols (Section 21.5)



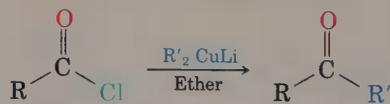
e. Partial reduction to yield aldehydes (Section 21.5)



f. Grignard reaction to yield tertiary alcohols (Section 21.5)

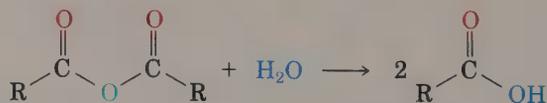


g. Diorganocopper reaction to yield ketones (Section 21.5)

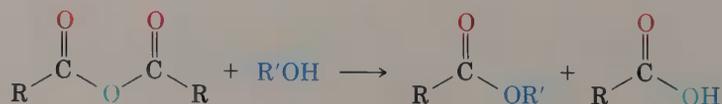


3. Reactions of acid anhydrides

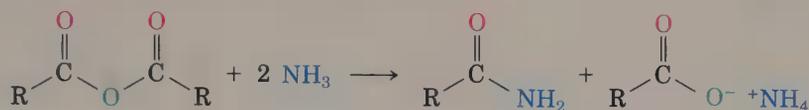
a. Hydrolysis to yield acids (Section 21.6)



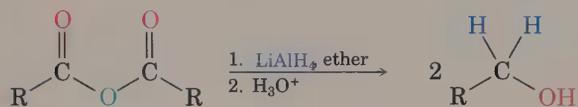
b. Alcoholysis to yield esters (Section 21.6)



c. Aminolysis to yield amides (Section 21.6)

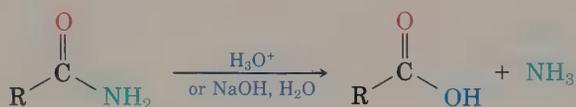


d. Reduction to yield primary alcohols (Section 21.6)

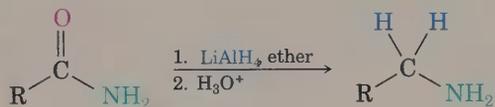


4. Reactions of amides and lactams

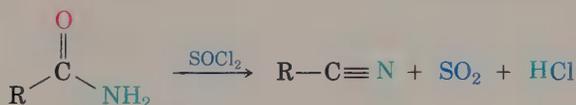
a. Hydrolysis to yield acids (Section 21.8)



b. Reduction to yield amines (Section 21.8)

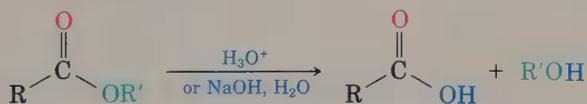


c. Dehydration of primary amides to yield nitriles (Section 21.9)

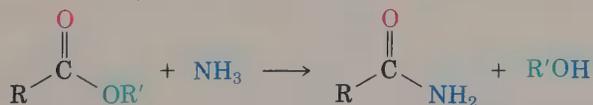


5. Reactions of esters and lactones

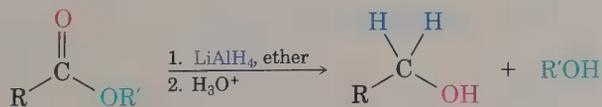
a. Hydrolysis to yield acids (Section 21.7)



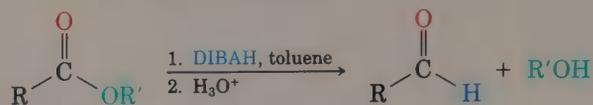
b. Aminolysis to yield amides (Section 21.7)



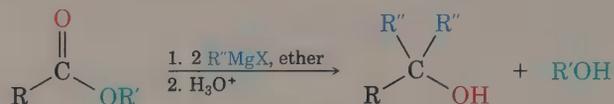
c. Reduction to yield primary alcohols (Section 21.7)



d. Partial reduction to yield aldehydes (Section 21.7)

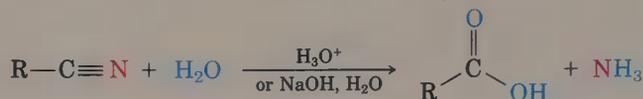


e. Grignard reaction to yield tertiary alcohols (Section 21.7)

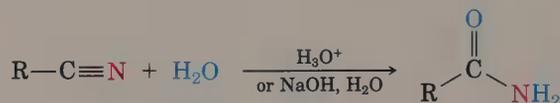


## 6. Reactions of nitriles

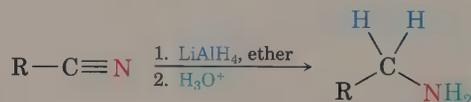
a. Hydrolysis to yield carboxylic acids (Section 21.9)



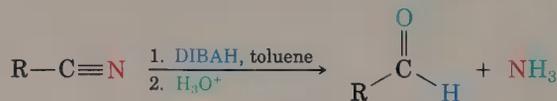
b. Partial hydrolysis to yield amides (Section 21.9)



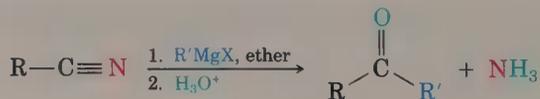
c. Reduction to yield primary amines (Section 21.9)



d. Partial reduction to yield aldehydes (Section 21.9)

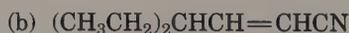
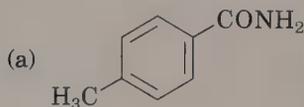


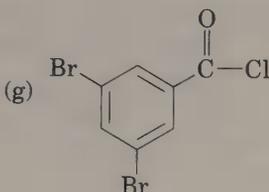
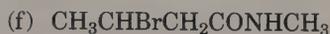
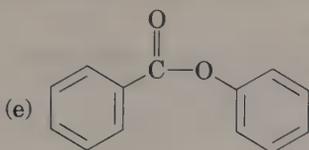
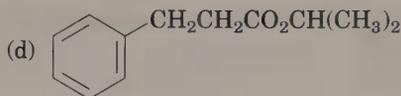
e. Reaction with Grignard reagents to yield ketones (Section 21.9)



## ADDITIONAL PROBLEMS .....

21.31 Give IUPAC names for these compounds:



**21.32** Draw structures corresponding to these names:

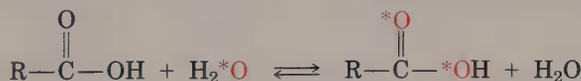
- (a) *p*-Bromophenylacetamide                      (b) *m*-Benzoylbenzotrile  
 (c) 2,2-Dimethylhexanamide                  (d) Cyclohexyl cyclohexanecarboxylate  
 (e) 2-Cyclobutenecarbonitrile                (f) 1,2-Pentanedicarbonyl dichloride

**21.33** Draw and name compounds that meet these descriptions:

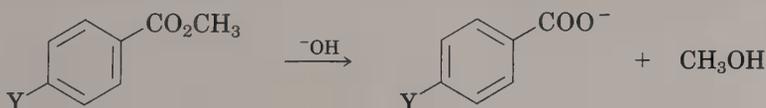
- (a) Three different acid chlorides having the formula  $\text{C}_6\text{H}_9\text{ClO}$   
 (b) Three different amides having the formula  $\text{C}_7\text{H}_{11}\text{NO}$   
 (c) Three different nitriles having the formula  $\text{C}_5\text{H}_7\text{N}$

**21.34** The following reactivity order has been found for the saponification of alkyl acetates by aqueous hydroxide ion:

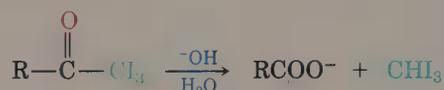
How can you explain this reactivity order?

**21.35** How can you explain the observation that attempted Fischer esterification of 2,4,6-trimethylbenzoic acid with methanol/HCl is unsuccessful? No ester is obtained, and the acid is recovered unchanged. Can you suggest an alternative method of esterification that would be successful?**21.36** How can you account for the fact that, when a carboxylic acid is dissolved in isotopically labeled water, the label rapidly becomes incorporated into *both* oxygen atoms of the carboxylic acid?**21.37** Outline methods for the preparation of acetophenone (phenyl methyl ketone) from the following:

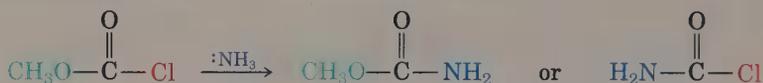
- (a) Benzene                                      (b) Bromobenzene                      (c) Methyl benzoate  
 (d) Benzotrile                                    (e) Styrene

**21.38** In the basic hydrolysis of *para*-substituted methyl benzoates, the following reactivity order has been found for Y:  $\text{NO}_2 > \text{Br} > \text{H} > \text{CH}_3 > \text{OCH}_3$ . How can you explain this reactivity order? Where would you expect  $\text{Y} = \text{C}\equiv\text{N}$ ,  $\text{Y} = \text{CHO}$ , and  $\text{Y} = \text{NH}_2$  to be in the reactivity list?

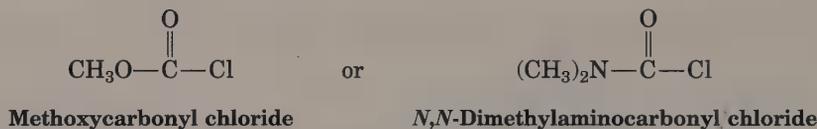
- 21.39 How might you prepare these compounds from butanoic acid?  
 (a) 1-Butanol (b) Butanal (c) 1-Bromobutane  
 (d) Pentanenitrile (e) 1-Butene (f) *N*-Methylpentanamide  
 (g) 2-Hexanone (h) Butylbenzene
- 21.40 What product would you expect to obtain from Grignard reaction of phenylmagnesium bromide with dimethyl carbonate,  $\text{CH}_3\text{OCOOCH}_3$ ?
- 21.41 When ethyl benzoate is heated in methanol containing a small amount of HCl, methyl benzoate is formed. Propose a mechanism for this reaction.
- 21.42 In the iodoform reaction, a triiodomethyl ketone reacts with aqueous base to yield a carboxylate ion and iodoform (triiodomethane). Propose a mechanism for this reaction.



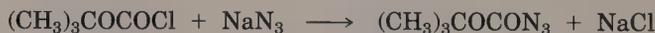
- 21.43 Which of the two products shown would you expect from the reaction of methoxycarbonyl chloride with ammonia? Would you expect the nucleophile to replace the chlorine or the methoxy group? Explain your answer.



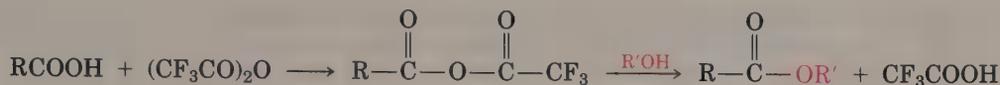
- 21.44 Which compound would you expect to have a more highly polarized carbonyl group, methoxycarbonyl chloride or *N,N*-dimethylaminocarbonyl chloride? Explain.



- 21.45 In light of your answer to Problem 21.44, which compound would you expect to react faster with nucleophiles: methoxycarbonyl chloride or *N,N*-dimethylaminocarbonyl chloride?
- 21.46 *tert*-Butoxycarbonyl azide, an important reagent used in protein synthesis, is prepared by treating *tert*-butoxycarbonyl chloride with sodium azide. Propose a mechanism for this reaction.



- 21.47 Predict the product, if any, of reaction between propanoyl chloride and the following reagents:  
 (a)  $(\text{Ph})_2\text{CuLi}$  in ether (b)  $\text{LiAlH}_4$ , then  $\text{H}_3\text{O}^+$   
 (c)  $\text{CH}_3\text{MgBr}$ , then  $\text{H}_3\text{O}^+$  (d)  $\text{Li}(\text{O}-t\text{-Bu})_3\text{AlH}$   
 (e)  $\text{H}_3\text{O}^+$  (f) Cyclohexanol  
 (g) Aniline (h)  $\text{CH}_3\text{COO}^- + \text{Na}$
- 21.48 Answer Problem 21.47 for reaction of the listed reagents with methyl propanoate.
- 21.49 Answer Problem 21.47 for reaction of the listed reagents with propanamide and with pentanenitrile.
- 21.50 A particularly mild method of esterification involves the use of trifluoroacetic anhydride. Treatment of a carboxylic acid with trifluoroacetic anhydride leads to a mixed anhydride that rapidly reacts with alcohol:



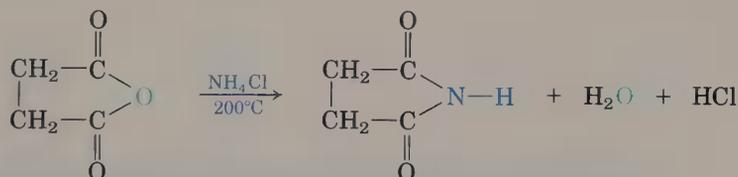
- Propose a mechanism for formation of the mixed anhydride.
- Why is the mixed anhydride unusually reactive?
- Why does the mixed anhydride react specifically as indicated, rather than giving trifluoroacetate esters plus carboxylic acid?

**21.51** How would you accomplish the following transformations? More than one step may be required.

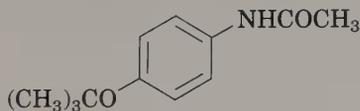
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CN} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{NH}_2$
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CN} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{N}(\text{CH}_3)_2$
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CN} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{C}(\text{CH}_3)_2\text{OH}$
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CN} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}(\text{OH})\text{CH}_3$
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CN} \longrightarrow \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CHO}$

**21.52** List as many ways as you can think of for transforming cyclohexanol into cyclohexanecarbaldehyde (try to get at least four).

**21.53** Succinic anhydride yields succinimide when heated with ammonium chloride at 200°C. Propose a mechanism for this reaction. Why do you suppose such a high reaction temperature is required?

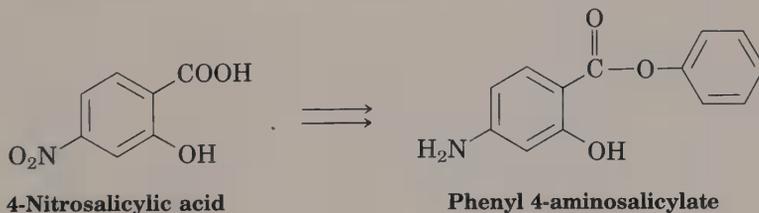


**21.54** Butacetin is an analgesic (pain killing) agent that is synthesized commercially from *p*-fluoronitrobenzene. Propose a likely synthesis route.



**Butacetin**

**21.55** Phenyl 4-aminosalicylate is a drug used in the treatment of tuberculosis. Propose a synthesis of this compound starting from 4-nitrosalicylic acid.



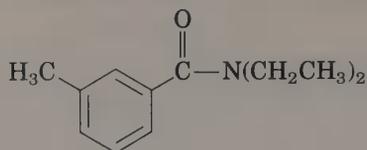
**4-Nitrosalicylic acid**

**Phenyl 4-aminosalicylate**

**21.56** What spectroscopic technique would you use to distinguish between the following isomer pairs? Tell what differences you would expect to see.

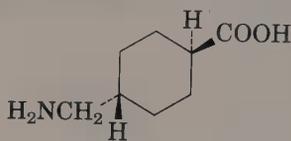
- N*-Methylpropanamide and *N,N*-dimethylacetamide
- 5-Hydroxypentanenitrile and cyclobutanecarboxamide
- 4-Chlorobutanoic acid and 3-methoxypropanoyl chloride
- Ethyl propanoate and propyl acetate

- 21.57 *N,N*-Diethyl-*m*-toluamide (DEET) is the active ingredient in many insect-repellent preparations. How might you synthesize this substance from *m*-bromotoluene?



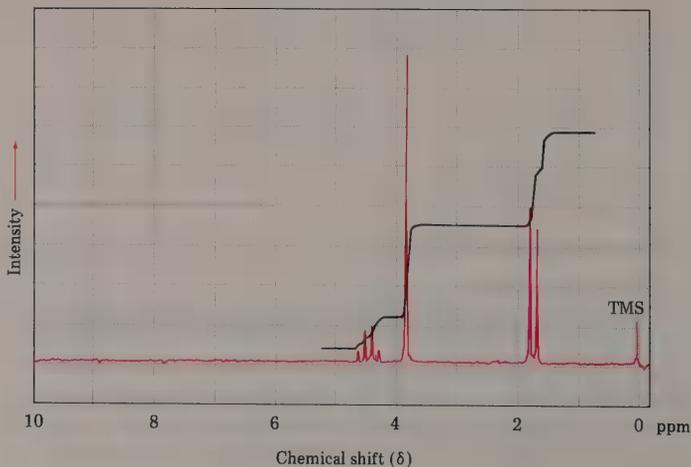
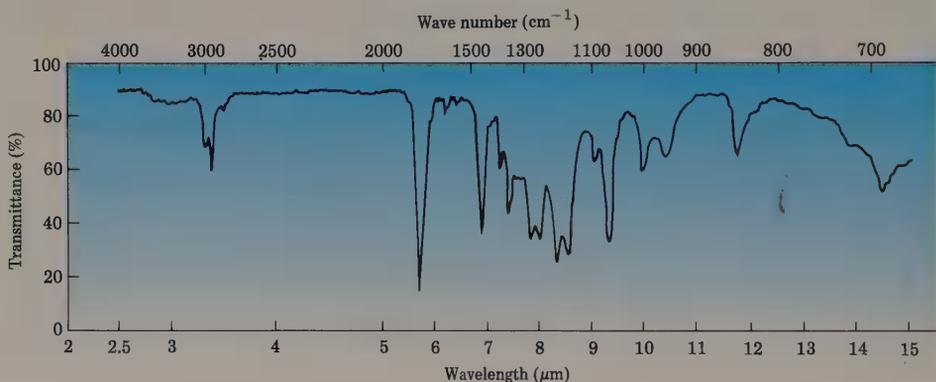
*N,N*-Diethyl-*m*-toluamide

- 21.58 Tranexamic acid, a drug useful against blood clotting, is prepared commercially from *p*-methylbenzotrile. Formulate the steps likely to be used in the synthesis. (Don't worry about *cis*-*trans* isomers. Heating to 300°C interconverts the isomers.)

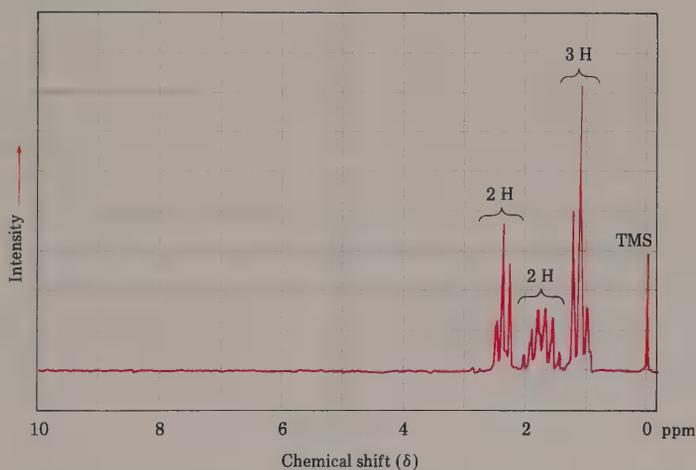
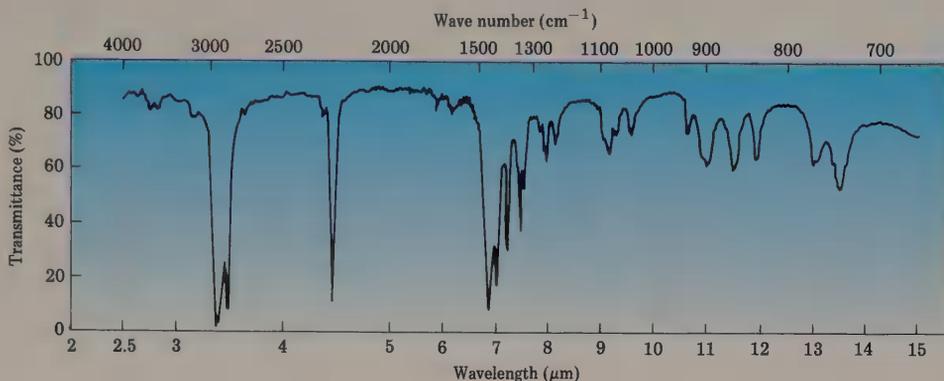


Tranexamic acid

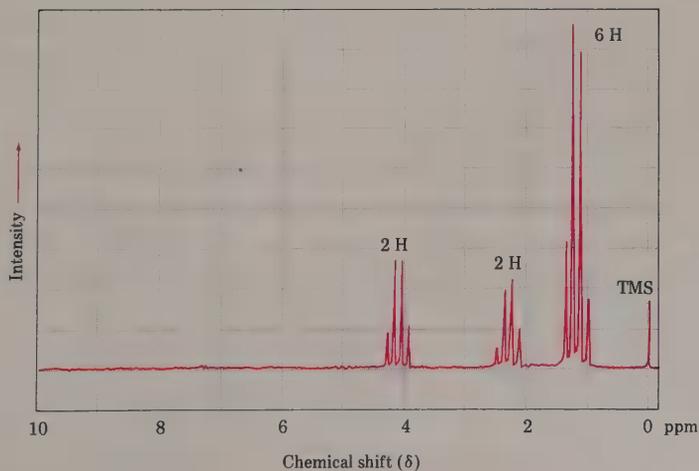
- 21.59 Propose a structure for a compound,  $C_4H_7ClO_2$ , that has the infrared and NMR spectra shown.



21.60 Propose a structure for a compound,  $C_4H_7N$ , that has the infrared and NMR spectra shown.

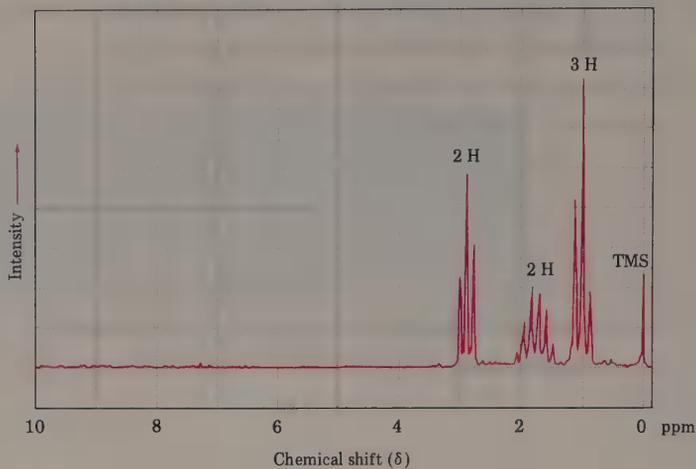


21.61 The following  $^1H$  NMR spectrum is of ethyl propanoate,  $CH_3CH_2COOCH_2CH_3$ . Assign all peaks in the spectrum, and account for the fact that the methyl groups appear as an evident *quartet* centered at 1.2  $\delta$ .

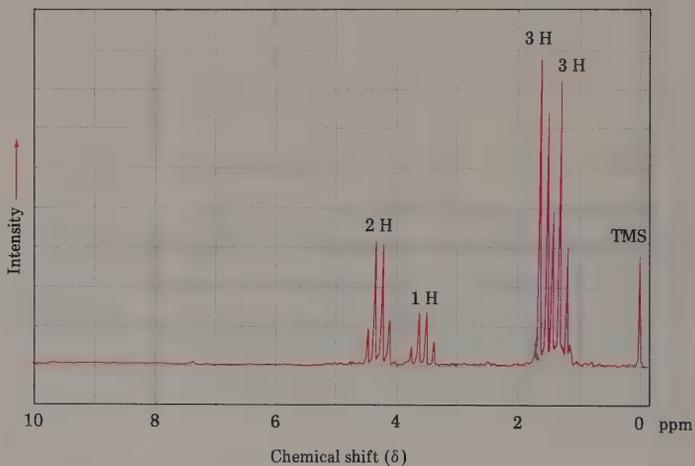


21.62 Assign structures to compounds with the following  $^1\text{H}$  NMR spectra:

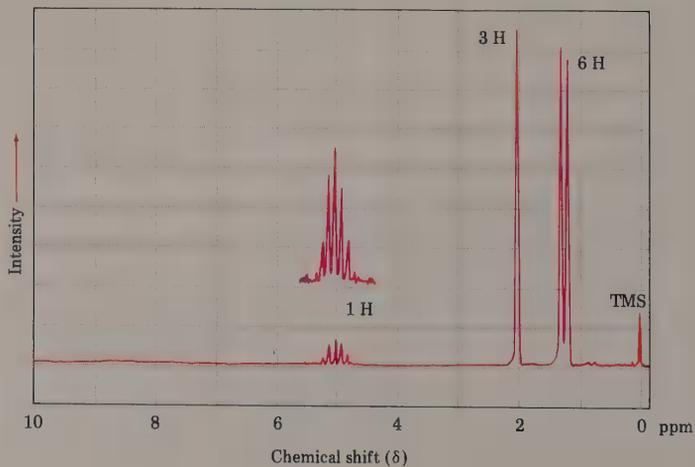
- (a)  $\text{C}_4\text{H}_7\text{ClO}$ ;  
IR:  $1810\text{ cm}^{-1}$



- (b)  $\text{C}_6\text{H}_9\text{NO}_2$ ;  
IR:  $2250$ ,  
 $1735\text{ cm}^{-1}$

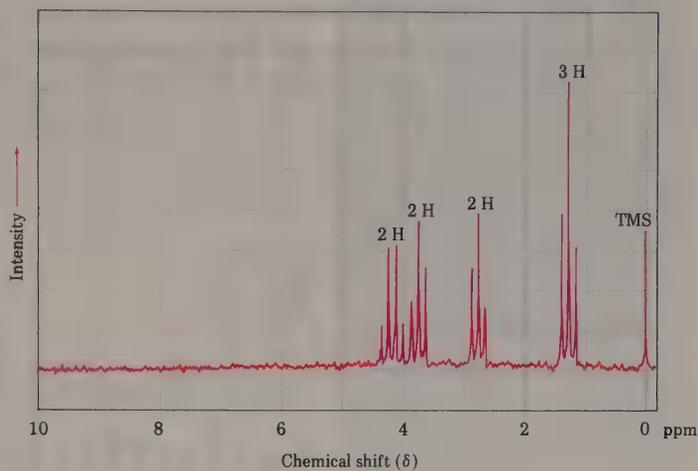


- (c)  $\text{C}_5\text{H}_{10}\text{O}_2$ ;  
IR:  $1735\text{ cm}^{-1}$

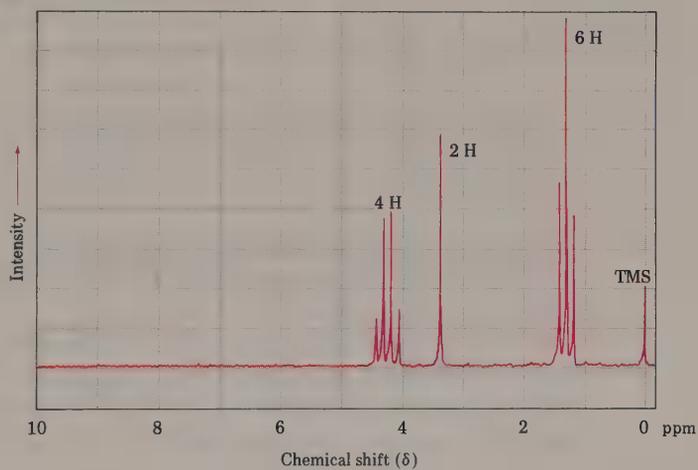


21.63 Propose structures for compounds with the following  $^1\text{H}$  NMR spectra:

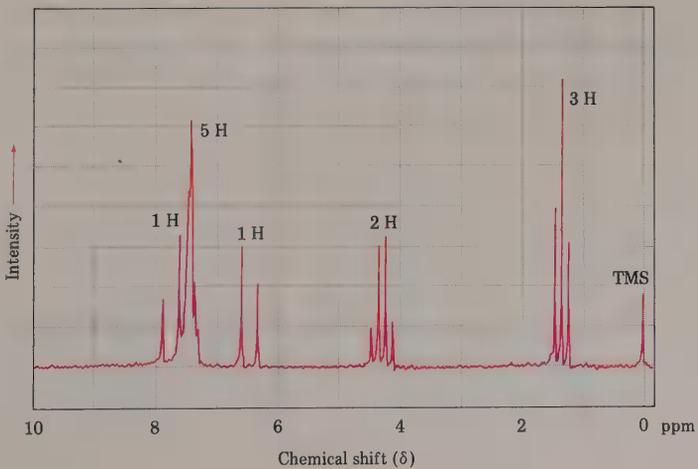
- (a)  $\text{C}_5\text{H}_9\text{ClO}_2$ ;  
IR:  $1735\text{ cm}^{-1}$

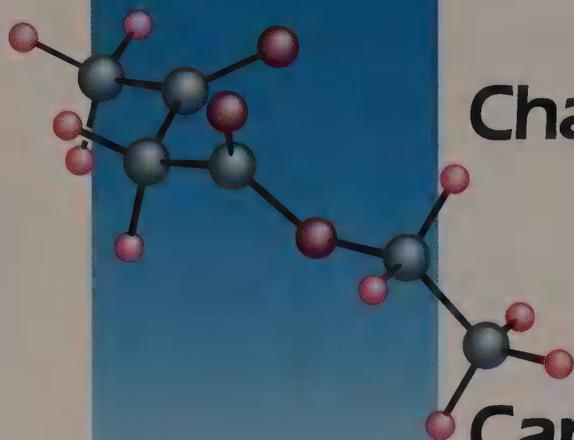


- (b)  $\text{C}_7\text{H}_{12}\text{O}_4$ ;  
IR:  $1735\text{ cm}^{-1}$



- (c)  $\text{C}_{11}\text{H}_{12}\text{O}_2$ ;  
IR:  $1710\text{ cm}^{-1}$



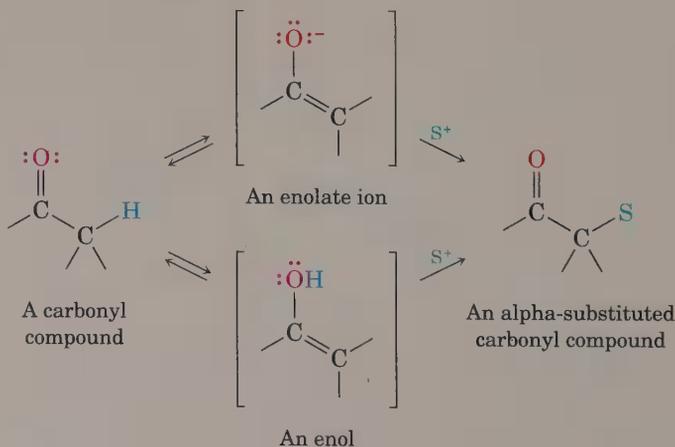


# Chapter 22

## Carbonyl Alpha-Substitution Reactions

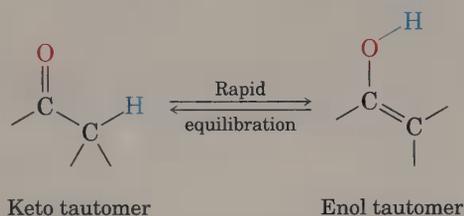
We said in the overview of carbonyl chemistry that most carbonyl-group chemistry can be explained in terms of just four fundamental reactions: nucleophilic additions, nucleophilic acyl substitutions, alpha substitutions, and carbonyl condensations. Having already looked at the chemistry of nucleophilic addition reactions and nucleophilic acyl substitution reactions, we'll now look at the chemistry of the third major carbonyl-group process—the **alpha-substitution reaction**.

Alpha-substitution reactions occur at the position *next to* the carbonyl group—the **alpha ( $\alpha$ ) position**—and involve the substitution of an alpha hydrogen atom by some other group. They take place through the formation of either enol or enolate ion intermediates. Let's begin our study by learning more about these two species.

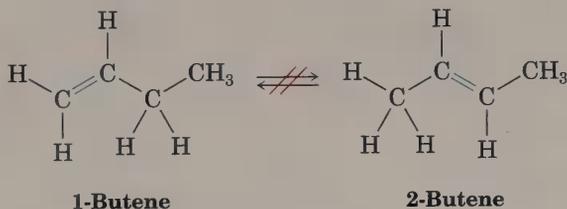


## 22.1 Keto–Enol Tautomerism

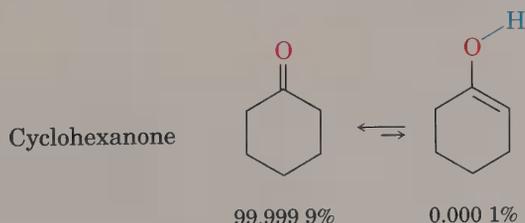
Carbonyl compounds that have hydrogen atoms on their alpha carbons are rapidly interconvertible with their corresponding **enols** (*ene* + *ol*, unsaturated alcohol). This rapid interconversion between two chemically distinct species is a special kind of isomerism known as **tautomerism** (taw-**tom**-erism; from the Greek *tauto*, “the same,” and *meros*, “part”). Individual isomers are called **tautomers** (taw-toe-mers).

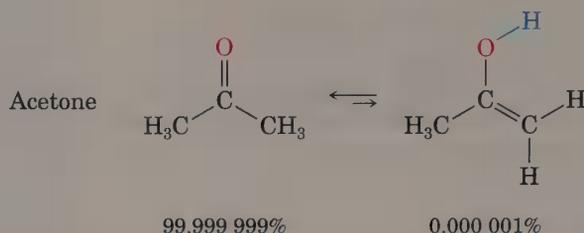


Note that tautomerism requires the two different isomeric forms to be *rapidly* interconvertible. Thus, keto and enol carbonyl isomers *are* tautomers, but two isomeric alkenes such as 1-butene and 2-butene are *not*, since they don't rapidly interconvert under normal circumstances.

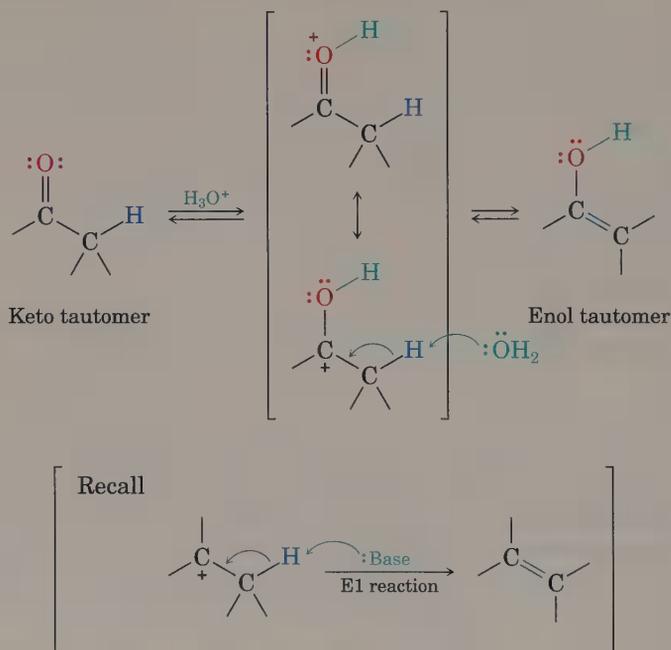


Most carbonyl compounds exist almost exclusively in the keto form at equilibrium, and it's usually difficult to isolate the pure enol form. For example, cyclohexanone contains only about 0.000 1% of its enol tautomer at room temperature, and acetone contains only about 0.000 001% enol. The percentage of enol tautomer is even less for carboxylic acids and acyl derivatives such as esters and amides. Even though enols are difficult to isolate and are present only to a small extent at equilibrium, they are nevertheless extremely important and are involved in much of the chemistry of carbonyl compounds.



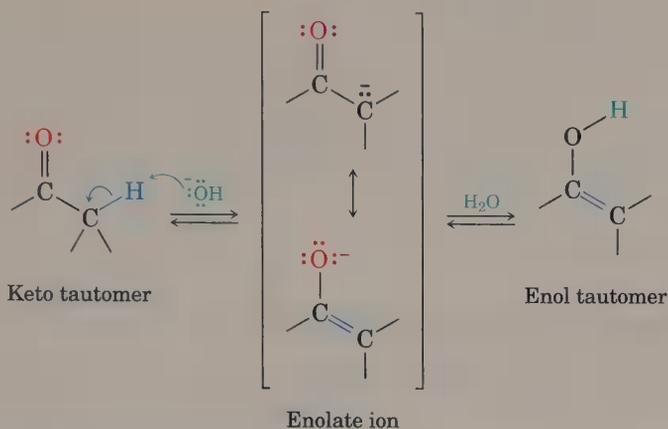


Keto–enol tautomerism of carbonyl compounds is catalyzed by both acids and bases. Acid catalysis involves protonation of the carbonyl oxygen atom (a Lewis base) to give a resonance-stabilized intermediate cation that can lose a proton from the alpha carbon to yield neutral enol (Figure 22.1). This proton loss from the cation intermediate is analogous to what occurs during an E1 reaction when a carbocation loses a proton to form an alkene (Section 11.14).



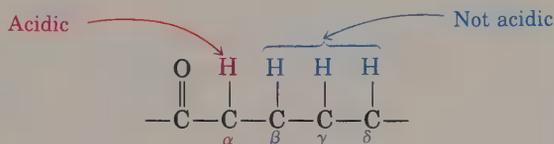
**Figure 22.1** Mechanism of acid-catalyzed enol formation.

Base-catalyzed enol formation occurs by an acid–base reaction between catalyst and carbonyl compound. The carbonyl compound acts as a weak protic acid and donates one of its alpha hydrogens to the base. The resultant anion—an **enolate ion**—is then reprotonated to yield a neutral compound. Since the enolate ion is a resonance hybrid of two forms, it can be protonated either on carbon to regenerate the keto tautomer or on oxygen to give an enol tautomer (Figure 22.2).



**Figure 22.2** Mechanism of base-catalyzed enol formation. The intermediate enolate anion, a resonance hybrid of two forms, can be protonated either on carbon to regenerate starting ketone or on oxygen to give an enol.

Note that only the protons on the *alpha* position of carbonyl compounds are acidic. Protons at beta, gamma, delta, and so on, are not acidic and can't be removed by base. We'll account for this unique behavior of alpha protons shortly.



PROBLEM.....

**22.1** Draw structures for the enol tautomers of these compounds:

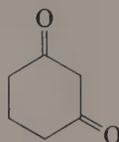
- |   |                     |                   |
|---|---------------------|-------------------|
| (a) Cyclopentanone                      | (b) Acetyl chloride | (c) Ethyl acetate |
| (d) Propanal                            | (e) Acetic acid     | (f) Phenylacetone |
| (g) Acetophenone (methyl phenyl ketone) |                     |                   |

PROBLEM.....

**22.2** How many acidic hydrogens does each of the molecules listed in Problem 22.1 have? Identify them.

PROBLEM.....

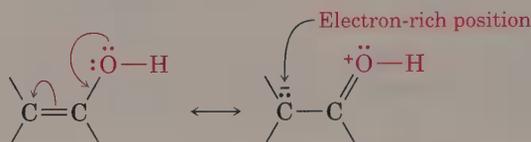
**22.3** Draw structures for the possible mono-enol forms of 1,3-cyclohexanedione. How many enol forms are possible? Which would you expect to be most stable? Explain your answer.



1,3-Cyclohexanedione

## 22.2 Reactivity of Enols: The Mechanism of Alpha-Substitution Reactions

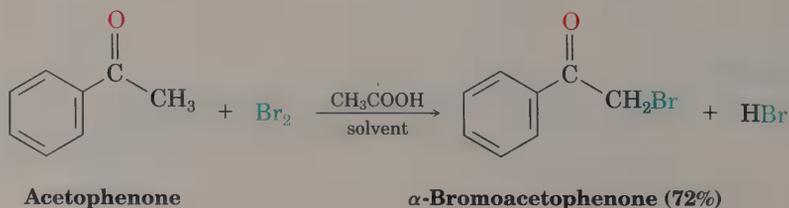
What kind of chemistry should enols have? Since their double bonds are electron-rich, enols behave as nucleophiles and react with electrophiles in much the same way that alkenes do. Because of resonance electron donation of the lone-pair electrons on oxygen, though, enol double bonds are more reactive than those of alkenes (Section 14.9).



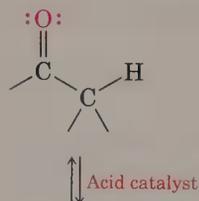
When an alkene reacts with an electrophile such as bromine, addition of  $\text{Br}^+$  occurs to give an intermediate cation, which reacts with  $\text{Br}^-$ . When an enol reacts with an electrophile, however, the initial addition step is the same but the intermediate cation loses the  $-\text{OH}$  proton to regenerate a carbonyl compound. The net result of the reaction of an enol with an electrophile is alpha substitution. The general mechanism is shown in Figure 22.3.

## 22.3 Alpha Halogenation of Ketones and Aldehydes

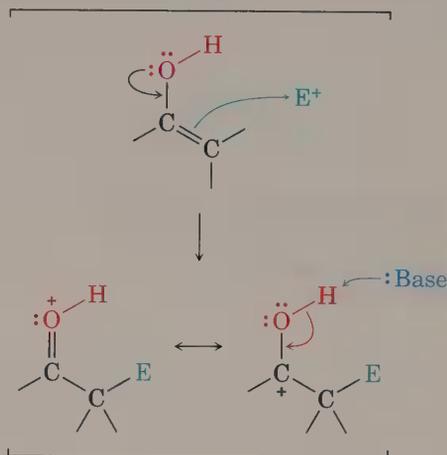
Ketones and aldehydes can be halogenated at their alpha positions by reaction with chlorine, bromine, or iodine in acidic solution. Bromine is most often used, and acetic acid is often employed as solvent.



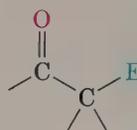
Acid-catalyzed enol formation occurs by the usual mechanism.



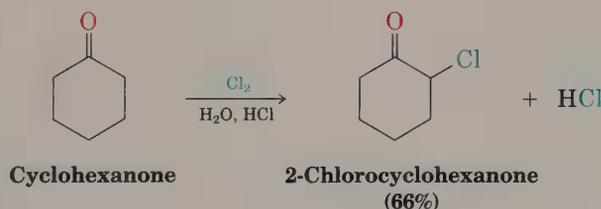
An electron pair from the enol oxygen attacks an electrophile, forming a new bond and leaving a cation intermediate that is stabilized by resonance between two forms.



Loss of a proton from oxygen yields the neutral alpha-substitution product as a new C=O bond is formed.



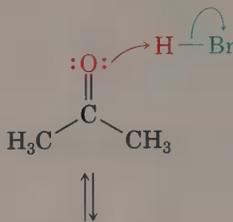
**Figure 22.3** General mechanism of a carbonyl  $\alpha$ -substitution reaction.



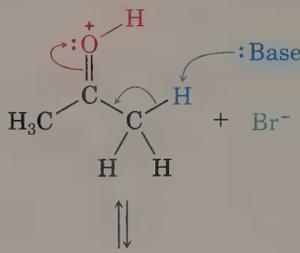
The alpha halogenation of ketones is a typical  $\alpha$ -substitution reaction that proceeds by acid-catalyzed formation of enol intermediates, as shown in Figure 22.4.

There is a great deal of evidence to indicate that the mechanism shown in Figure 22.4 is correct. For example, the rate of the halogenation reaction is independent of the nature of the halogen. Chlorination, bromination, and iodination of a given ketone all occur at the same rate, indicating that the same rate-limiting step is involved in chlorination, bromination, and iodination.

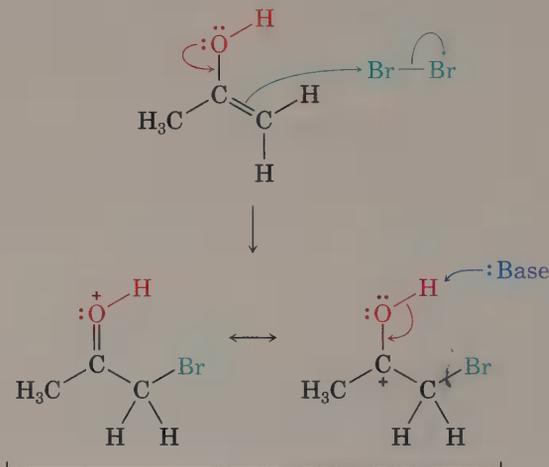
The carbonyl oxygen atom is protonated by acid catalyst.



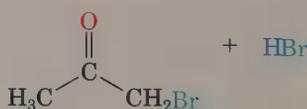
Loss of an acidic proton from the alpha carbon takes place in the normal way to yield an enol intermediate.



An electron pair from the enol attacks bromine, giving an intermediate cation that is stabilized by resonance between two forms.



Loss of the -OH proton then gives the alpha-halogenated product and generates more acid catalyst.



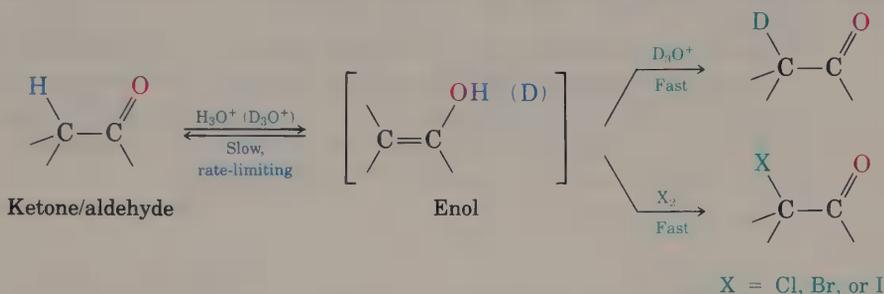
**Figure 22.4** Mechanism of the acid-catalyzed bromination of ketones and aldehydes.

Additional evidence comes from measurements indicating that acid-catalyzed ketone halogenations exhibit second-order kinetics and follow the rate law:

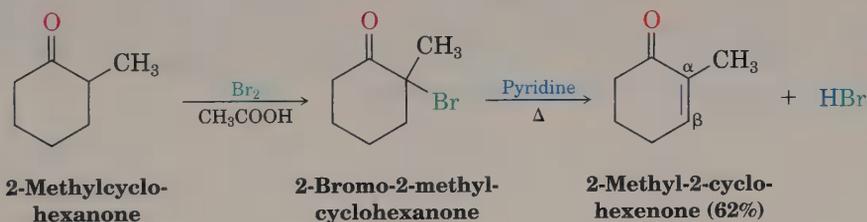
$$\text{Reaction rate} = k[\text{Ketone}][\text{H}^+]$$

This information tells us that ketone halogenations depend only on ketone and acid concentrations but are independent of halogen concentration. Thus, halogen is not involved in the rate-limiting step.

A final piece of evidence comes from deuteration experiments. If a ketone is treated with  $D_3O^+$  instead of  $H_3O^+$ , the acidic  $\alpha$ -hydrogen atoms are replaced by deuterium atoms. For a given ketone, the rate of deuteration is identical to the rate of halogenation, indicating that the same intermediate is involved in both deuteration and halogenation. The common intermediate that satisfies all the evidence can only be an enol.

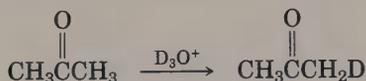


$\alpha$ -Bromo ketones are useful in organic synthesis because they can be dehydrobrominated by base treatment to yield  $\alpha,\beta$ -unsaturated ketones. For example, 2-bromo-2-methylcyclohexanone gives 2-methyl-2-cyclohexenone in 62% yield when refluxed in pyridine. The reaction, which takes place by an E2 elimination pathway (Section 11.11), is an excellent method for introducing carbon-carbon double bonds into molecules.



PROBLEM .....

22.4 Show in detail the mechanism of the deuteration of acetone on treatment with  $D_3O^+$ .



PROBLEM .....

22.5 When optically active (*R*)-3-phenyl-2-butanone is exposed to aqueous acid, a loss of optical activity occurs and racemic 3-phenyl-2-butanone is produced. Explain.

PROBLEM .....

22.6 In light of your answer to Problem 22.5, would you expect optically active (*R*)-3-methyl-3-phenyl-2-pentanone to be racemized by acid treatment? Explain.

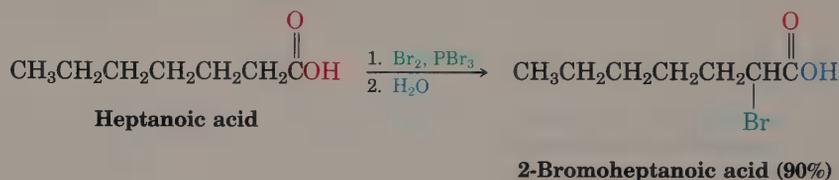
## PROBLEM .....

22.7 Show how you might prepare 1-penten-3-one from 3-pentanone.

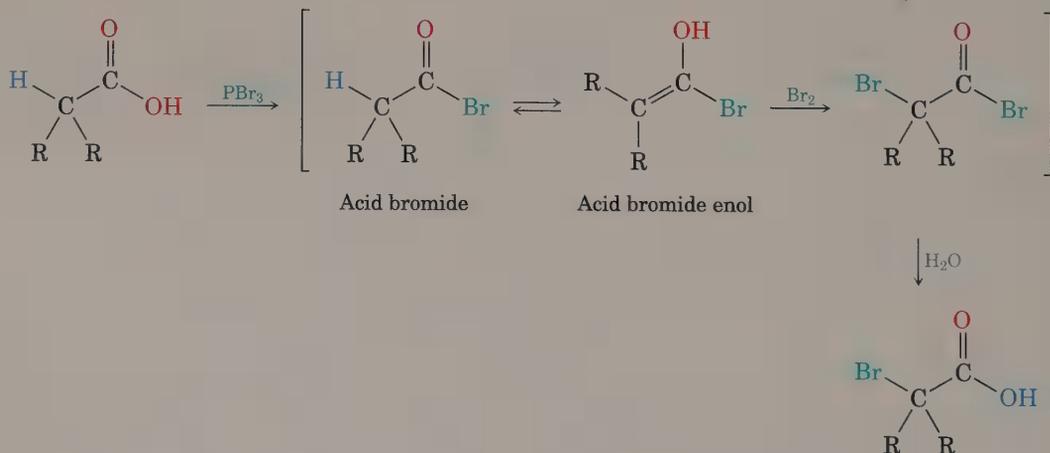
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## 22.4 Alpha Bromination of Carboxylic Acids: The Hell–Volhard–Zelinskii Reaction

Direct alpha bromination of carbonyl compounds by bromine in acetic acid is limited to ketones and aldehydes, since acids, esters, and amides don't enolize sufficiently for halogenation to take place. Carboxylic acids, however, can be  $\alpha$ -brominated by a mixture of bromine and phosphorus tribromide—the **Hell–Volhard–Zelinskii (HVZ) reaction**:



The first step in the Hell–Volhard–Zelinskii reaction takes place between  $\text{PBr}_3$  and carboxylic acid to yield an intermediate acid bromide plus  $\text{HBr}$  (Section 21.5). The  $\text{HBr}$  produced in the first step next catalyzes enolization of the acid bromide, and the enol reacts rapidly with bromine. Hydrolysis of the  $\alpha$ -bromo acid bromide by addition of water then gives the  $\alpha$ -bromo carboxylic acid product.



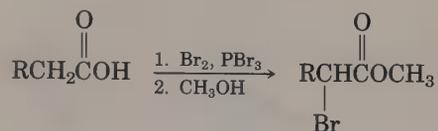
The overall result of the Hell–Volhard–Zelinskii reaction is the transformation of an acid into an  $\alpha$ -bromo acid. Note, though, that the key step involves bromination of an acid bromide enol in a manner analogous to that occurring during ketone bromination.

## PROBLEM

- 22.8 If an optically active carboxylic acid such as (*R*)-2-phenylpropanoic acid were brominated under Hell–Volhard–Zelinskii conditions, would you expect the product to be optically active or racemic? Explain.

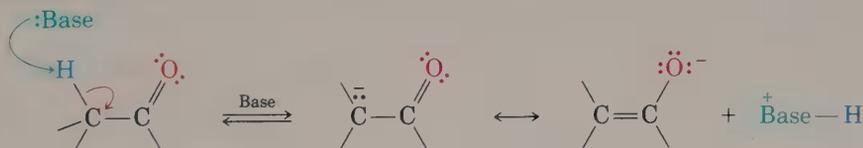
## PROBLEM

- 22.9 If methanol rather than water is added at the end of the Hell–Volhard–Zelinskii reaction, an ester rather than an acid is produced. Propose a mechanism for this transformation.

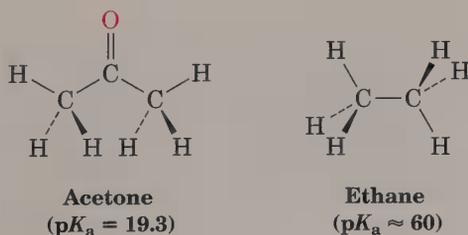


## 22.5 Acidity of Alpha-Hydrogen Atoms: Enolate Ion Formation

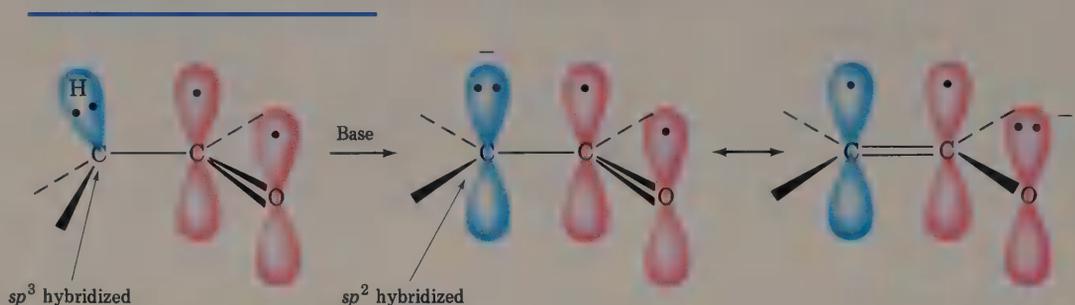
During the discussion of base-catalyzed enol formation earlier in this chapter (Section 22.1), we said that carbonyl compounds act as weak protic acids. Strong bases can therefore abstract acidic alpha protons from carbonyl compounds to yield enolate anions:



Why are carbonyl compounds weakly acidic? If we compare acetone ( $\text{p}K_{\text{a}} \approx 19.3$ ) with ethane ( $\text{p}K_{\text{a}} \approx 60$ ), we find that the presence of a neighboring carbonyl group increases the acidity of a ketone over an alkane by a factor of  $10^{40}$ .

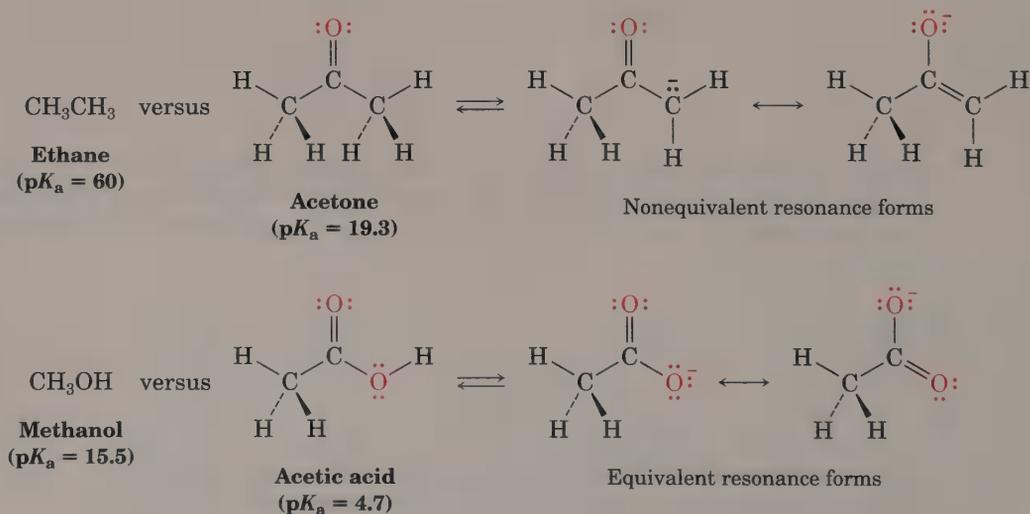


The reason for this increased acidity is best seen by looking at an orbital picture of the enolate ion (Figure 22.5). Proton abstraction from a carbonyl compound occurs when the alpha C–H sigma bond is oriented in the plane of the carbonyl-group  $p$  orbitals. The alpha-carbon atom of the enolate ion product is  $sp^2$  hybridized and has a  $p$  orbital that overlaps the neighboring carbonyl-group  $p$  orbitals. Thus, the negative charge is shared by the electronegative oxygen atom, and the enolate ion is stabilized by resonance between two forms.

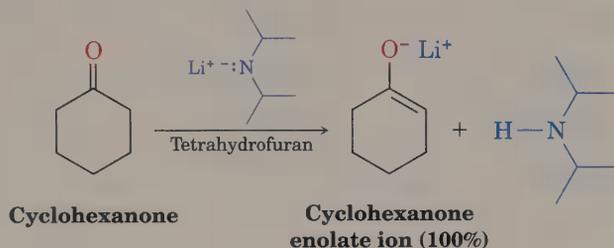


**Figure 22.5** Mechanism of enolate ion formation by abstraction of an alpha proton from a carbonyl compound.

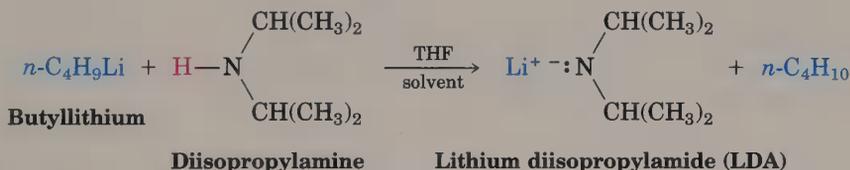
Carbonyl compounds are more acidic than alkanes for the same reason that carboxylic acids are more acidic than alcohols (Section 20.3). In both cases, the anions are stabilized by resonance. Enolate ions differ from carboxylate ions in that their two resonance forms are not equivalent: The form with the negative charge on oxygen is lower in energy than the form with the charge on carbon. Nevertheless, the principle behind resonance stabilization is the same in both cases.



Since alpha-hydrogen atoms of carbonyl compounds are only weakly acidic when compared with carboxylic acids, strong bases must be used to effect enolate ion formation. If an alkoxide ion such as sodium ethoxide is used as base, ionization takes place only to the extent of about 0.1%, because ethanol ( $pK_a = 16$ ) is a stronger acid than acetone is. If, however, a more powerful base such as sodium hydride (NaH, the sodium "salt" of  $H_2$ ), sodium amide ( $NaNH_2$ , the sodium salt of ammonia), or lithium diisopropylamide [ $LiN(i-C_3H_7)_2$ , the lithium salt of diisopropylamine] is used, a carbonyl compound can be completely converted into its enolate ion.



Lithium diisopropylamide, usually abbreviated LDA, is easily prepared by reaction between butyllithium and diisopropylamine and is widely used as a base for preparing enolate ions from carbonyl compounds. LDA has nearly ideal properties: It's an exceedingly powerful base because diisopropylamine has  $pK_a \approx 40$ ; it's soluble in organic solvents such as THF; it doesn't add to carbonyl groups in nucleophilic addition reactions; and it's extremely reactive, even at  $-78^\circ C$ .

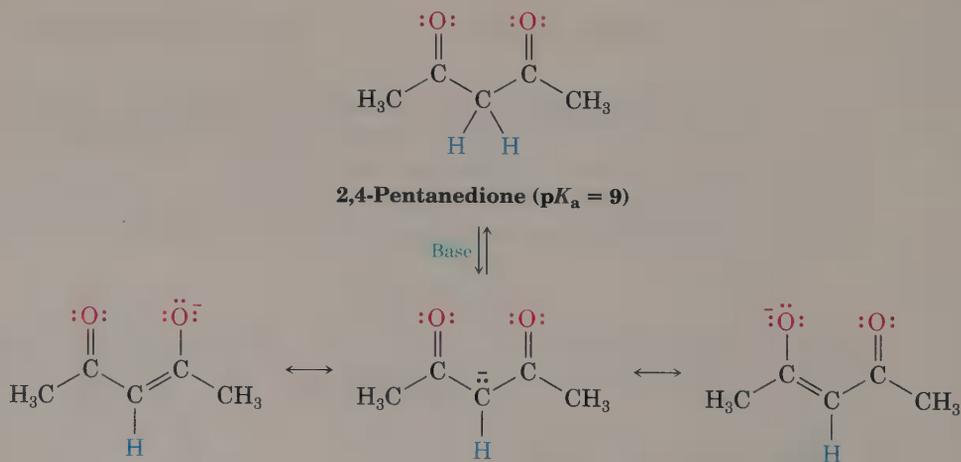


All types of carbonyl compounds, including aldehydes, ketones, esters, acid chlorides, and amides, have greatly enhanced  $\alpha$ -hydrogen acidity compared to alkanes. Thus, all can be converted into enolate ions by reaction with LDA. Table 22.1 lists the approximate  $pK_a$  values of different kinds of carbonyl compounds and shows how these values compare with other acidic substances we've seen.

When a hydrogen atom is flanked by two carbonyl groups, its acidity is enhanced even more. Thus, Table 22.1 shows that such compounds as 1,3-diketones ( **$\beta$ -diketones**), 1,3-keto esters ( **$\beta$ -keto esters**), and 1,3-diester are more acidic than water is. This enhanced acidity of  $\beta$ -dicarbonyl compounds is due to the fact that the resultant enolate ions are highly stabilized by resonance delocalization of the negative charge over *both* neighboring carbonyl groups. For example, the enolate ion of 2,4-pentanedione has three resonance forms:

TABLE 22.1 Acidity Constants for Some Organic Compounds

Compound type	Compound	p <i>K</i> <sub>a</sub>
Carboxylic acid	CH <sub>3</sub> COOH	5
1,3-Diketone	CH <sub>2</sub> (COCH <sub>3</sub> ) <sub>2</sub>	9
1,3-Keto ester	CH <sub>3</sub> COCH <sub>2</sub> CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	11
1,3-Dinitrile	CH <sub>2</sub> (CN) <sub>2</sub>	11
1,3-Diester	CH <sub>2</sub> (CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub> ) <sub>2</sub>	13
Water	HOH	16
Primary alcohol	CH <sub>3</sub> CH <sub>2</sub> OH	16
Acid chloride	CH <sub>3</sub> COCl	16
Aldehyde	CH <sub>3</sub> CHO	17
Ketone	CH <sub>3</sub> COCH <sub>3</sub>	19
Ester	CH <sub>3</sub> CO <sub>2</sub> C <sub>2</sub> H <sub>5</sub>	25
Nitrile	CH <sub>3</sub> CN	25
Dialkylamide	CH <sub>3</sub> CON(CH <sub>3</sub> ) <sub>2</sub>	30
Ammonia	NH <sub>3</sub>	35
Dialkylamine	HN( <i>i</i> -C <sub>3</sub> H <sub>7</sub> ) <sub>2</sub>	40
Alkyne	HC≡CH	25
Alkene	CH <sub>2</sub> =CH <sub>2</sub>	49
Alkane	CH <sub>3</sub> CH <sub>3</sub>	60



Similar resonance forms can be drawn for other doubly stabilized enolate ions.

PROBLEM .....

22.10 Identify the most acidic hydrogens in these molecules:

- (a)  $\text{CH}_3\text{CH}_2\text{CHO}$                       (b)  $(\text{CH}_3)_3\text{CCOCH}_3$                       (c)  $\text{CH}_3\text{COOH}$   
 (d) Benzamide                              (e)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CN}$                       (f)  $\text{CH}_3\text{CON}(\text{CH}_3)_2$   
 (g) 1,3-Cyclohexanedione

PROBLEM .....

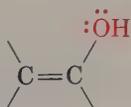
22.11 When optically active (*R*)-2-methylcyclohexanone is treated with aqueous NaOH, racemic 2-methylcyclohexanone is produced. Propose a mechanism to account for this racemization.

PROBLEM .....

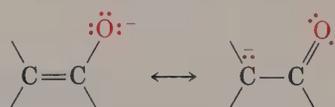
22.12 Would you expect optically active (*S*)-3-methylcyclohexanone to be racemized on base treatment in the same way that 2-methylcyclohexanone is (Problem 22.11)? Explain.

## 22.6 Reactivity of Enolate Ions

Enolate ions are more useful than enols for two reasons. First, pure enols can't normally be isolated. They are usually generated only as transient intermediates in low concentration. By contrast, stable solutions of pure enolate ions are easily prepared from most carbonyl compounds by reaction with a strong base. Second and more important, enolate ions are much more reactive than enols; they undergo many important reactions that enols don't. Whereas enols are neutral, enolate ions carry a full negative charge, making them much better nucleophiles. Thus, the  $\alpha$ -carbon atom of an enolate ion is highly reactive toward electrophiles.

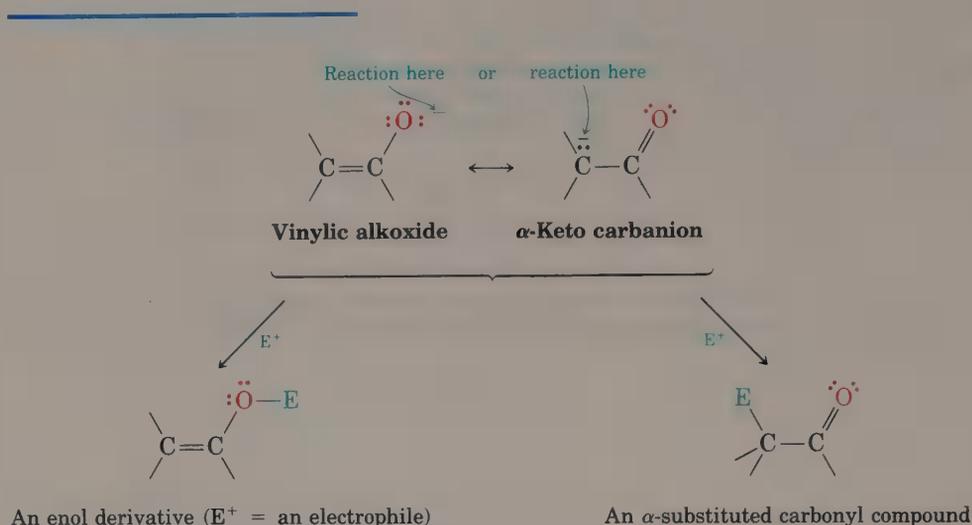


Enol: neutral, moderately reactive,  
very difficult to isolate



Enolate: negatively charged, very reactive,  
easily prepared

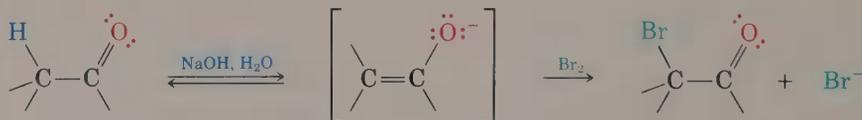
Since enolate ions are resonance hybrids of two nonequivalent resonance forms, they can be looked at either as vinylic alkoxides ( $\text{C}=\text{C}-\text{O}^-$ ) or as  $\alpha$ -keto carbanions ( $^-\text{C}-\text{C}=\text{O}$ ). Thus, enolate ions can react with electrophiles either on oxygen or on carbon. Reaction on oxygen yields an enol derivative, whereas reaction on carbon yields an  $\alpha$ -substituted carbonyl compound (Figure 22.6). Both kinds of reactivity are known, but reaction on carbon is the more commonly observed pathway.



**Figure 22.6** Two modes of enolate ion reactivity. Reaction on carbon to yield an  $\alpha$ -substituted carbonyl product is the more commonly followed path.

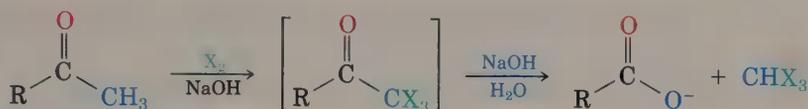
## 22.7 Halogenation of Enolate Ions: The Haloform Reaction

Halogenation of ketones is promoted by base as well as by acid. As you might expect, the base-promoted reaction occurs through enolate ion intermediates. Even relatively weak bases such as hydroxide ion are effective for halogenation, since it's not necessary to convert the ketone completely into its enolate ion. It's only necessary to generate a small amount of enolate at any one time because the reaction with halogen occurs as soon as the enolate ion is formed.



Base-promoted halogenation of ketones is little used in practice because it's difficult to stop the reaction at the monosubstituted product. An  $\alpha$ -halogenated ketone product is generally more acidic than the starting unsubstituted ketone because of the electron-withdrawing inductive effect of the halogen atom. Thus, the monohalo products are themselves rapidly turned into enolates and further halogenated.

If excess base and halogen are used, methyl ketones are triply halogenated and then cleaved by base in the **haloform reaction**:



A methyl ketone

where X = Cl, Br, or I

The haloform reaction, which converts a methyl ketone into a carboxylic acid plus a **haloform** (chloroform,  $\text{CHCl}_3$ ; bromoform,  $\text{CHBr}_3$ ; or iodoform,  $\text{CHI}_3$ ), is the basis for a qualitative test for methyl ketones. A sample of unknown structure is dissolved in THF or ether and placed in a test tube. Dilute solutions of aqueous sodium hydroxide and iodine are then added, and the test tube is observed. Formation of a yellow precipitate of iodoform,  $\text{CHI}_3$ , signals a positive test and indicates that the sample is a methyl ketone.

PROBLEM .....

- 22.13 Base-promoted chlorination and bromination of a given ketone occur at the same rate. Explain.

PROBLEM .....

- 22.14 Why do you suppose that ketone halogenations in acidic media are referred to as being acid *catalyzed*, whereas halogenations in basic media are base *promoted*? Why is a full equivalent of base required for halogenation?

PROBLEM .....

- 22.15 The second step of the haloform reaction is simply a nucleophilic acyl substitution: the replacement of  $-\text{CX}_3$  by  $-\text{OH}$ . Show the mechanism of this step.
- .....

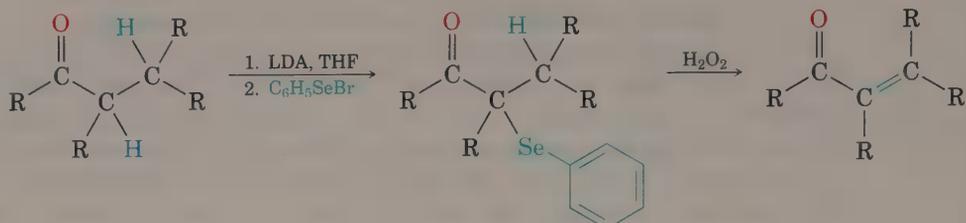
## 22.8 Selenenylation of Enolate Ions: Enone Synthesis

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One of the most important recent advances in enolate ion chemistry occurred in the mid-1970s when it was found that carbonyl compounds can be *selenenylated*. That is, a selenium atom can be attached to the alpha position of a carbonyl compound. Selenenylation is accomplished by first allowing the carbonyl compound to react with a strong base such as LDA in THF solvent, followed by treatment of the enolate ion solution with one equivalent of benzeneselenenyl bromide ( $\text{C}_6\text{H}_5\text{SeBr}$ ). Immediate  $\alpha$ -substitution reaction then yields an  $\alpha$ -phenylseleno-substituted product.

The great value of the selenenylation reaction lies in the fact that the product can be easily converted into an  $\alpha,\beta$ -unsaturated carbonyl compound.

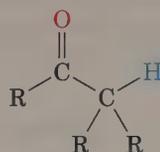
On treatment with dilute hydrogen peroxide at room temperature, the selenium group is oxidized, elimination occurs, and an  $\alpha,\beta$ -unsaturated carbonyl compound is formed. The net result is introduction of a carbon-carbon double bond into the  $\alpha,\beta$  position of the saturated carbonyl starting material. Yields are usually excellent, and the method is often superior to the alternative  $\alpha$ -bromination-dehydrobromination route.



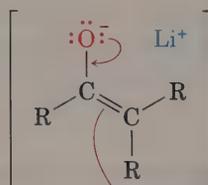
An  $\alpha$ -phenylseleno ketone

Mechanistically, the  $\alpha$ -phenylselenenylation reaction is similar to halogenation except that the reactive enolate ion intermediate is prepared in a separate step prior to reaction, rather than during reaction as in halogenation. The overall process is shown in Figure 22.7.

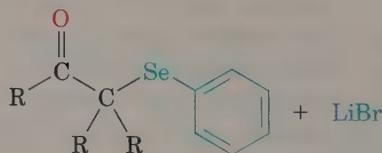
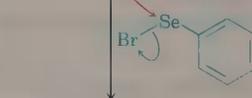
The strong base LDA abstracts an acidic alpha proton from the carbonyl compound, yielding an enolate ion.



LDA

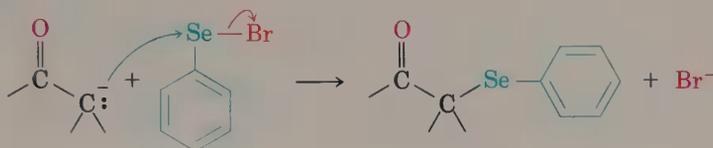


The nucleophilic enolate ion attacks selenium and carries out a displacement of bromide ion, yielding the  $\alpha$ -phenylseleno product.

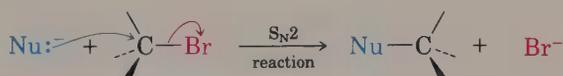


**Figure 22.7** Mechanism of the reaction of an enolate ion with benzeneselenenyl bromide.

Note that the key step in the alpha phenylselenenylation of carbonyl compounds is similar to an  $S_N2$  reaction (Section 11.4). The alpha-carbon atom of the enolate ion is the nucleophile that displaces bromide by attack on selenium:



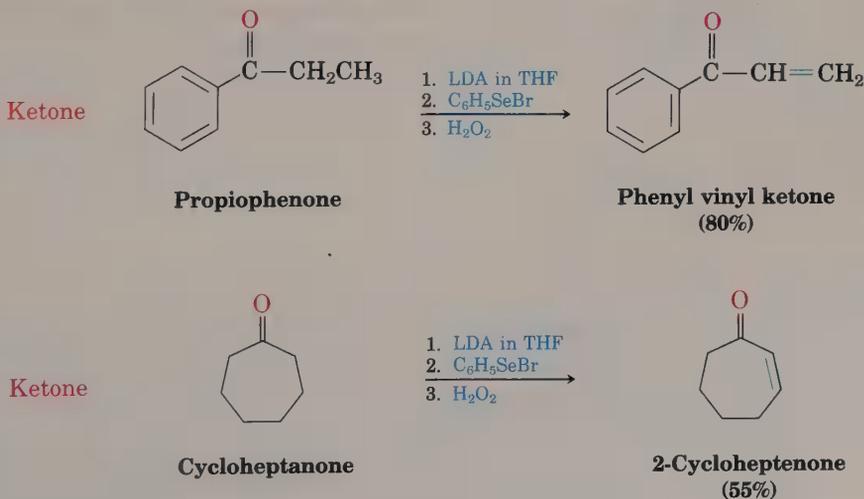
Recall

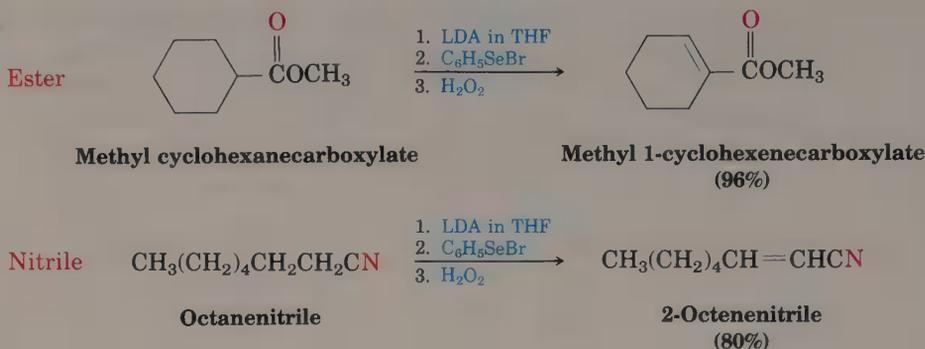


The elimination step involves oxidation of the phenylseleno intermediate to a phenyl selenoxide, which undergoes a spontaneous intramolecular elimination reaction. No added base is required (as in dehydrobromination), and the reaction occurs at room temperature under very mild conditions.



Ketones, esters, and nitriles all undergo the  $\alpha$ -phenylselenenylation-elimination reaction in good yield. Aldehydes, however, often give product mixtures because their enolate ions are too reactive.





## PROBLEM .....

**22.16** Show how you would prepare the following unsaturated molecules from saturated precursors.

(a) 1-Penten-3-one

(b) 4-*tert*-Butyl-2-cyclohexenone

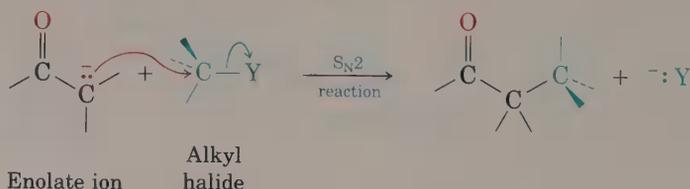
(c) 2-Methyl-2-pentenenitrile

## PROBLEM .....

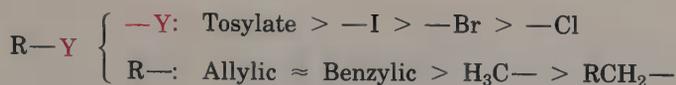
**22.17** Ketones react slowly with benzeneselenenyl chloride in the presence of HCl to yield  $\alpha$ -phenylseleno ketones. Propose a mechanism for this acid-catalyzed  $\alpha$ -substitution reaction. To what other reaction is this process analogous?

## 22.9 Alkylation of Enolate Ions

One of the most important reactions of enolate ions is their **alkylation** by treatment with an alkyl halide. The alkylation reaction is extremely useful because it allows the formation of a new carbon-carbon bond, thereby joining two small pieces into one larger molecule. Alkylation occurs when the nucleophilic enolate anion reacts with the electrophilic alkyl halide in an  $S_N2$  reaction and displaces the leaving group by back-side attack.

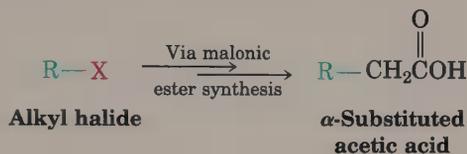


Alkylation reactions are subject to the same constraints that affect all  $S_N2$  reactions (Section 11.4). Thus, the leaving group,  $-Y$ , in the alkylating agent can be chloride, bromide, iodide, or tosylate. The alkyl group,  $R$ , must be primary or methyl and preferably should be allylic or benzylic. Secondary halides react poorly, and tertiary halides don't alkylate at all since a competing  $E2$  elimination of  $HY$  occurs instead. Vinylic and aryl halides are also unreactive, because back-side attack is sterically prevented.

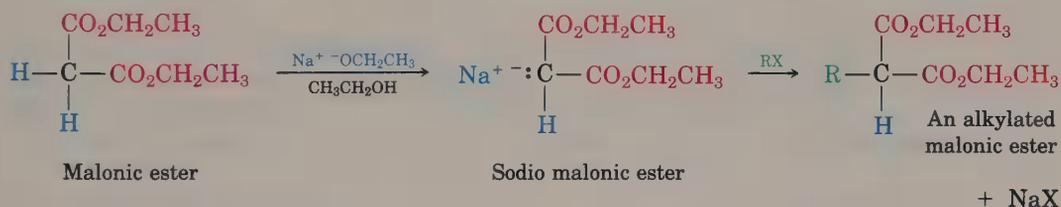


## THE MALONIC ESTER SYNTHESIS

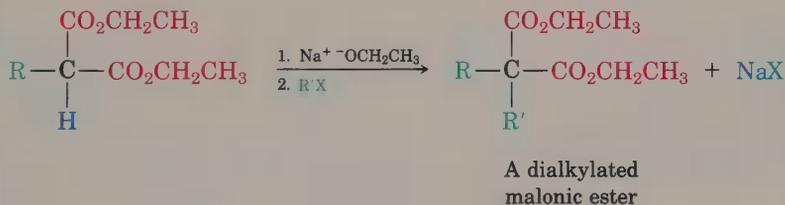
The **malonic ester synthesis**, one of the oldest and best-known carbonyl alkylation reactions, is an excellent method for preparing substituted acetic acids from alkyl halides:



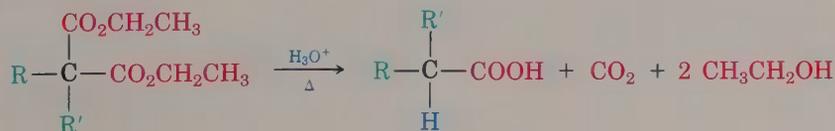
Diethyl propanedioate, commonly called diethyl malonate or *malonic ester*, is more acidic than many other carbonyl compounds ( $pK_a = 13$ ) because its  $\alpha$ -hydrogen atoms are flanked by two carbonyl groups. Thus, malonic ester is easily converted into its enolate anion by reaction with sodium ethoxide in ethanol. The enolate ion, in turn, is a good nucleophile that reacts rapidly with alkyl halides, yielding  $\alpha$ -substituted malonic esters.



The product of malonic ester alkylation has one acidic  $\alpha$ -hydrogen atom left, and the alkylation process can therefore be repeated a second time to yield a dialkylated malonic ester:

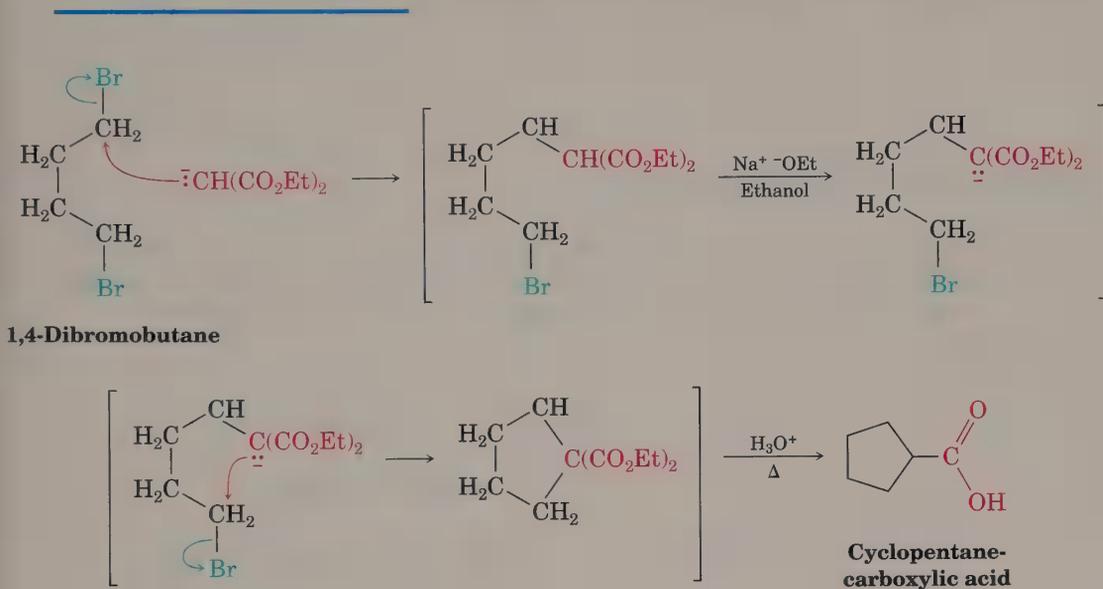


Once formed, the alkylated malonic esters can be hydrolyzed and decarboxylated when heated with aqueous acid (recall that decarboxylation is the loss of carbon dioxide,  $CO_2$ ). The product is a substituted monoacid.





The malonic ester synthesis can also be used to prepare *cycloalkane-carboxylic acids* by the proper choice of alkyl halide. For example, if 1,4-dibromobutane is treated with diethyl malonate in the presence of two equivalents of sodium ethoxide base, the first alkylation occurs as expected, but the second alkylation step occurs *internally* to yield a five-membered-ring product. Hydrolysis and decarboxylation then lead to cyclopentanecarboxylic acid (Figure 22.8). Three-, four-, five-, and six-membered rings can be prepared in this manner, but yields decrease for larger ring sizes.

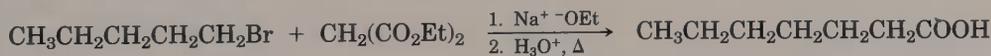


**Figure 22.8** Malonic ester synthesis of cyclopentanecarboxylic acid. The abbreviation “Et” means an ethyl group.

#### PRACTICE PROBLEM.....

How would you prepare heptanoic acid via a malonic ester synthesis?

**Solution** The malonic ester synthesis converts an alkyl halide into a carboxylic acid having two more carbons in its chain. Thus, a seven-carbon acid chain must be derived from the five-carbon alkyl halide 1-bromopentane.



#### PROBLEM.....

**22.18** What alkyl halide would you use to prepare these compounds by a malonic ester synthesis?

(a) Butanoic acid

(b) 5-Methylhexanoic acid

## PROBLEM

22.19 How could you use a malonic ester synthesis to prepare these compounds? Show all steps.

- (a) 3-Phenylpropanoic acid  
 (c) 4-Methylpentanoic acid

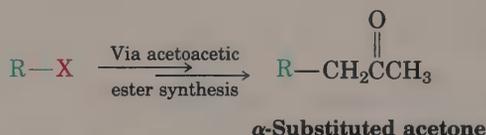
- (b) 2-Methylpentanoic acid  
 (d) Ethyl cyclobutanecarboxylate

## PROBLEM

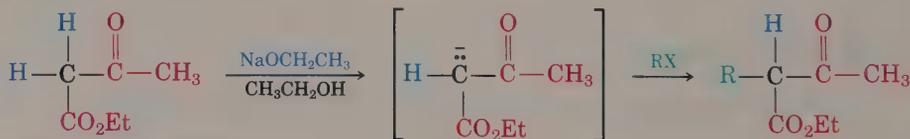
22.20 Monoalkylated and dialkylated acetic acids can be prepared by the malonic ester synthesis, but trialkylated acetic acids ( $R_3CCOOH$ ) can't be prepared. Explain.

## THE ACETOACETIC ESTER SYNTHESIS

The **acetoacetic ester synthesis** is a method for preparing  $\alpha$ -substituted acetone derivatives from alkyl halides in the same way that the malonic ester synthesis is a method for preparing  $\alpha$ -substituted acetic acids.

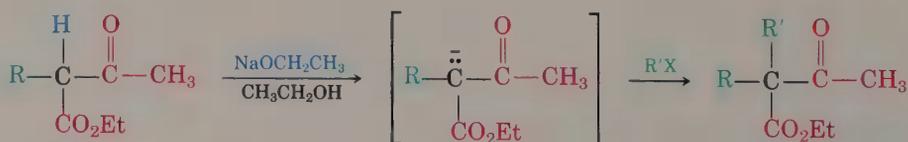


Ethyl 3-oxobutanoate, commonly called ethyl acetoacetate or *acetoacetic ester*, is much like malonic ester in that its alpha hydrogens are flanked by two carbonyl groups. It is therefore readily converted into its enolate ion, which can be alkylated by reaction with an alkyl halide. A second alkylation can also be carried out, if desired, since acetoacetic ester has two acidic alpha protons that can be replaced.



Acetoacetic ester

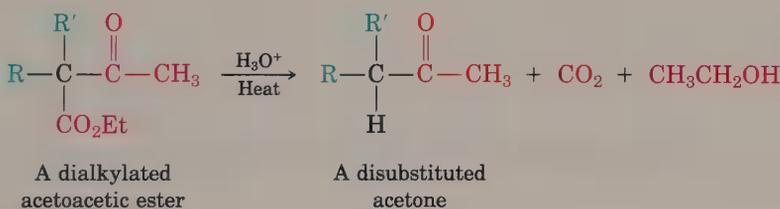
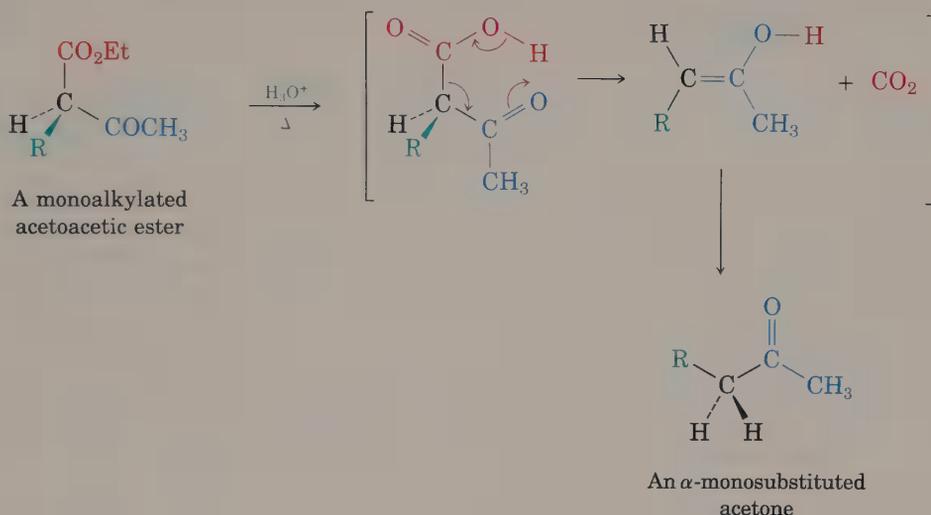
A monoalkylated acetoacetic ester



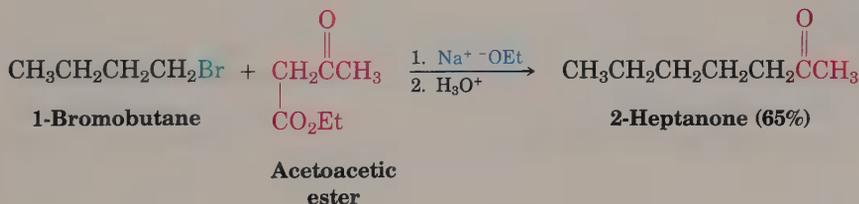
A monoalkylated acetoacetic ester

A dialkylated acetoacetic ester

On heating with aqueous HCl, the alkylated acetoacetic ester is hydrolyzed and decarboxylated via a  $\beta$ -keto acid intermediate to yield an  $\alpha$ -substituted acetone product. If a monoalkylated acetoacetic ester is hydrolyzed and decarboxylated, an  $\alpha$ -monosubstituted acetone is formed; if a dialkylated acetoacetic ester is hydrolyzed and decarboxylated, an  $\alpha,\alpha$ -disubstituted acetone is formed.



For example:



The three-step sequence of (1) enolate ion formation, (2) alkylation, and (3) hydrolysis–decarboxylation is applicable to all  $\beta$ -keto esters with acidic  $\alpha$  hydrogens, not just to acetoacetic ester itself. Thus, cyclic  $\beta$ -keto esters such as ethyl 2-oxocyclohexanecarboxylate can be alkylated and decarboxylated to give 2-substituted cyclohexanones in high yield. For example:



## PROBLEM

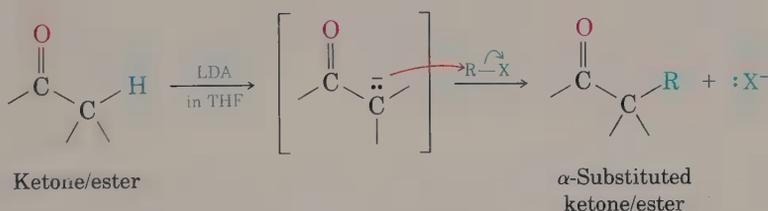
22.23 Which of the following compounds can't be prepared by an acetoacetic ester synthesis? Explain.

- (a) 2-Butanone (b) Phenylacetone  
 (c) Acetophenone (d) 3,3-Dimethyl-2-butanone

## DIRECT ALKYLATION OF KETONES, ESTERS, AND NITRILES

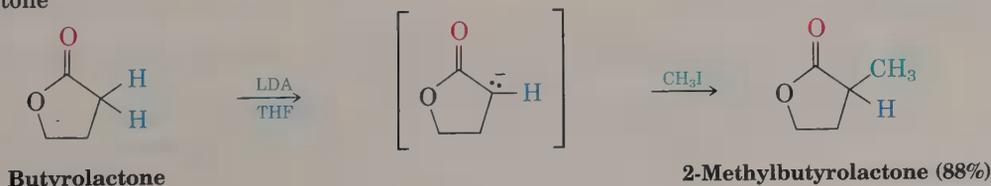
The malonic ester synthesis and the acetoacetic ester synthesis are special, since both alkylation reactions take place at doubly carbonyl-activated centers. By contrast, it's also possible in certain cases to alkylate at the alpha position of *monoketones*, *monoesters*, and nitriles.

The experimental conditions necessary to carry out the direct alkylation of carbonyl compounds are precise. Solvent and reaction temperature are both important, and the exact nature of the base used to generate the enolate ion is critical. The base must be sufficiently strong to convert a carbonyl compound ( $pK_a \approx 20\text{--}25$ ) into its enolate anion quickly, yet it must also be sufficiently bulky that it won't add to the carbonyl group in a nucleophilic addition or substitution reaction. Research carried out in the 1970s showed that LDA in THF solvent is highly effective in promoting alkylation reactions of carbonyl compounds.

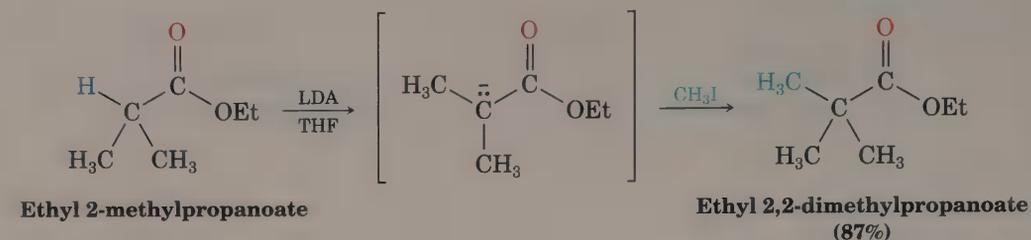


Ketones, esters, and nitriles can all be alkylated by using LDA or related dialkylamide bases in THF, but aldehydes rarely give high yields of pure products. (The problem is not that aldehydes aren't acidic enough; rather, it's that aldehyde enolate ions undergo carbonyl condensation reactions instead of alkylation. We'll study this condensation reaction in the next chapter.) Some specific examples of alkylation reactions are shown.

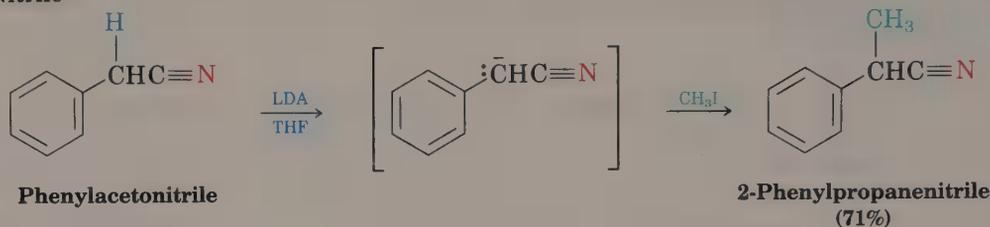
## Lactone



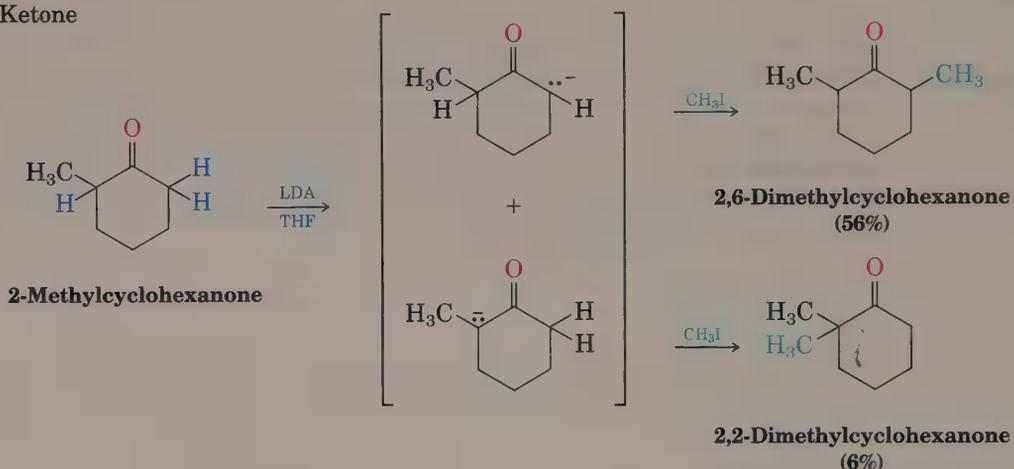
Ester



Nitrile



Ketone



Note in the previous examples that alkylation of the unsymmetrically substituted ketone 2-methylcyclohexanone leads to a mixture of products, since both possible enolate ions are formed. In general, the major product in such cases is derived from alkylation at the less hindered, more accessible position. Thus, alkylation of 2-methylcyclohexanone occurs primarily at the secondary 6-position, rather than at the tertiary 2-position.

## PROBLEM .....

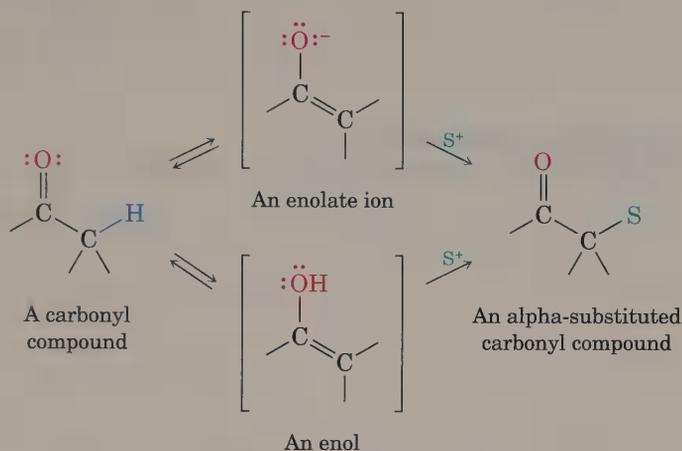
**22.24** How might you prepare the following compounds, using an alkylation reaction as the key step?

- (a) 3-Phenyl-2-butanone  
(c) 2-Allylcyclohexanone

- (b) 2-Ethylpentanenitrile  
(d) 2,2,6,6-Tetramethylcyclohexanone
- .....

## 22.10 Summary and Key Words

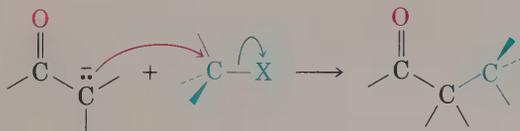
The **alpha substitution** of carbonyl compounds via **enol** or **enolate ion** intermediates is one of the four fundamental reaction types in carbonyl-group chemistry.



All carbonyl compounds rapidly equilibrate with their enols, a process called **tautomerism**. Although enol tautomers are normally present to only a small extent at equilibrium and can't usually be isolated in pure form, they nevertheless contain a highly nucleophilic double bond and react rapidly with a variety of electrophiles. For example, ketones and aldehydes are rapidly halogenated at the alpha position by reaction with chlorine, bromine, or iodine in acetic acid solution. Alpha bromination of carboxylic acids can be similarly accomplished by the **Hell-Volhard-Zelinskii reaction**, in which an acid is treated with  $\text{Br}_2/\text{PBr}_3$ . The  $\alpha$ -halogenated products can then undergo base-induced  $\text{E}2$  elimination to yield  $\alpha,\beta$ -unsaturated carbonyl products.

Alpha-hydrogen atoms of carbonyl compounds are acidic and can be abstracted by strong bases to yield enolate ions. Ketones, aldehydes, esters, amides, and nitriles can all be deprotonated in this manner if a sufficiently powerful base such as **lithium diisopropylamide (LDA)** is used. Enolate ions are highly reactive as nucleophiles because of their negative charge. For example, enolates react with benzeneselenenyl bromide to yield  $\alpha$ -phenylselenenylated products. These, in turn, yield  $\alpha,\beta$ -unsaturated carbonyl compounds when treated with  $\text{H}_2\text{O}_2$ .

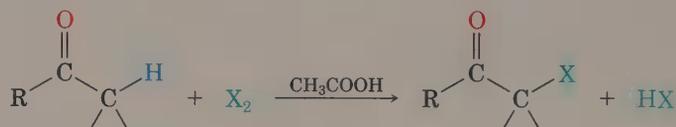
The most important reaction of enolates is their  $\text{S}_{\text{N}}2$  **alkylation** by alkyl halides. The nucleophilic enolate ion attacks an alkyl halide from the back side and displaces the leaving halide group to yield an alpha-alkylated carbonyl product.



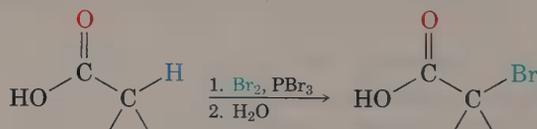
The **malonic ester synthesis**, which involves alkylation of diethyl malonate with an alkyl halide, provides a method for preparing monoalkylated or dialkylated acetic acids. Similarly, the **acetoacetic ester synthesis** provides a method for preparing monoalkylated or dialkylated acetone derivatives. Most important of all is that many carbonyl compounds, including ketones, esters, and nitriles, can be directly alkylated. These carbonyl alkylation reactions constitute what is perhaps the most important method in organic chemistry for synthesizing complex molecules.

## 22.11 Summary of Reactions

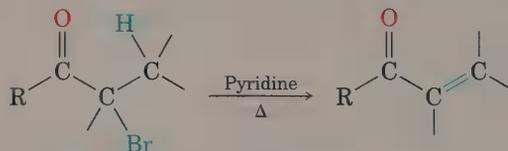
1. Ketone/aldehyde halogenation, where X = Cl, Br, or I (Section 22.3)



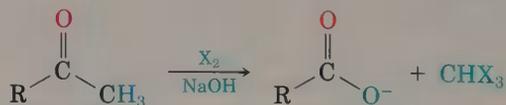
2. Hell-Volhard-Zelinskii bromination of acids (Section 22.4)



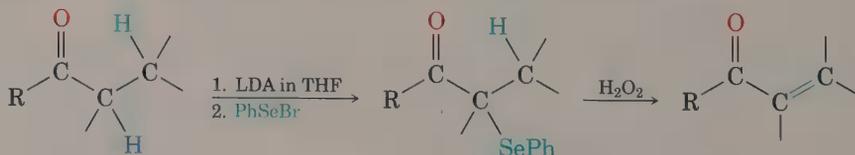
3. Dehydrobromination of  $\alpha$ -bromo ketones (Section 22.3)



4. Haloform reaction, where X = Cl, Br, or I (Section 22.7)

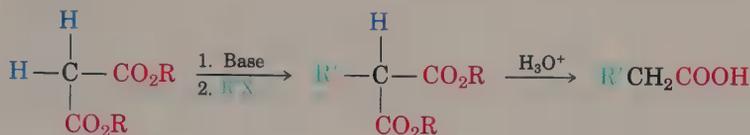


5. Phenylselenenylation-elimination (Section 22.8)

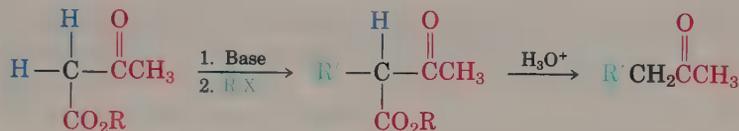


## 6. Alkylation of enolate ions (Section 22.9)

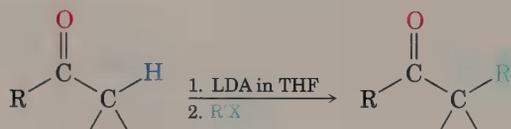
## a. Malonic ester synthesis



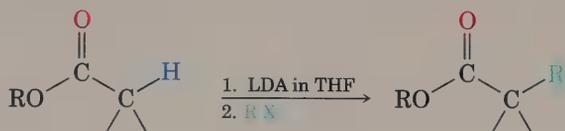
## b. Acetoacetic ester synthesis



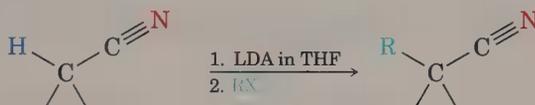
## c. Alkylation of ketones



## d. Alkylation of esters



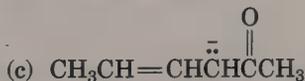
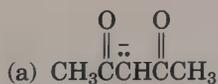
## e. Alkylation of nitriles

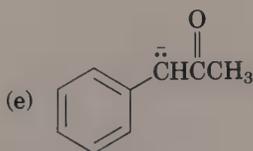


## ADDITIONAL PROBLEMS .....

22.25 Acetone is enolized only to the extent of about 0.000 01% at equilibrium, whereas 2,4-pentanedione is 76% enolized. Explain.

22.26 Write resonance structures for these anions:





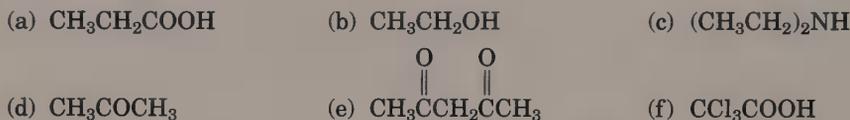
22.27 Indicate all the acidic hydrogen atoms in the following structures:



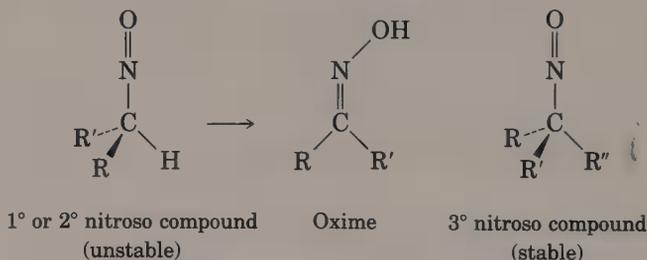
22.28 One way to determine the number of acidic hydrogens in a molecule is to treat the compound with NaOD in  $\text{D}_2\text{O}$ , isolate the product, and determine its molecular weight by mass spectroscopy. For example, if cyclohexanone is treated with NaOD in  $\text{D}_2\text{O}$ , the product has a molecular weight of 102. Explain how this method works.

22.29 2-Methylcycloheptanone and 3-methylcycloheptanone are nearly indistinguishable by spectroscopic techniques. How could you use the method outlined in Problem 22.28 to differentiate them?

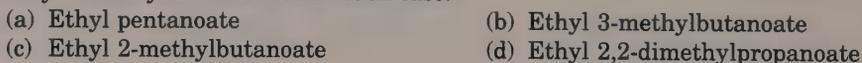
22.30 Rank the following compounds in order of increasing acidity:



22.31 All attempts to isolate primary and secondary nitroso compounds result only in the formation of oximes. Tertiary nitroso compounds, however, are stable. Explain.



22.32 Which of these compounds can be prepared by a malonic ester synthesis? Show the alkyl halide you would use in each case.



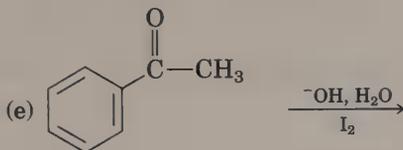
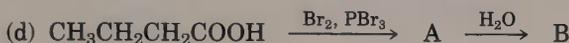
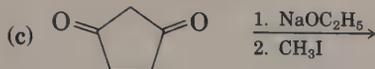
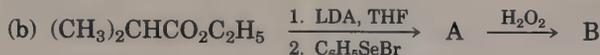
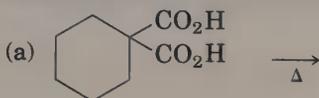
22.33 How would you prepare these ketones using an acetoacetic ester synthesis?



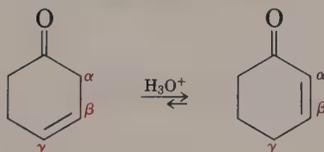
22.34 How would you prepare these compounds using either an acetoacetic ester synthesis or a malonic ester synthesis?



22.35 Predict the product(s) of these reactions:

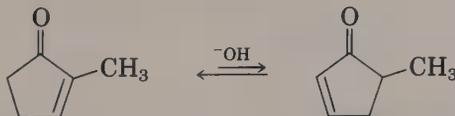


22.36 Nonconjugated  $\beta,\gamma$ -unsaturated ketones such as 3-cyclohexenone are in an acid-catalyzed equilibrium with their conjugated  $\alpha,\beta$ -unsaturated isomers. Propose a mechanism for the acid-catalyzed isomerization.

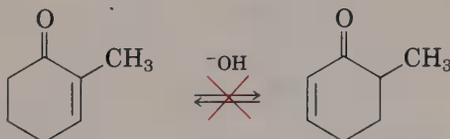


22.37 The  $\beta,\gamma$ -to- $\alpha,\beta$  interconversion of unsaturated ketones (Problem 22.36) is also catalyzed by base. Explain.

22.38 One interesting consequence of the base-catalyzed  $\beta,\gamma$ -to- $\alpha,\beta$  isomerization of unsaturated ketones (Problem 22.37) is that 2-substituted 2-cyclopentenones can be interconverted with 5-substituted 2-cyclopentenones. Propose a mechanism to account for this isomerization.

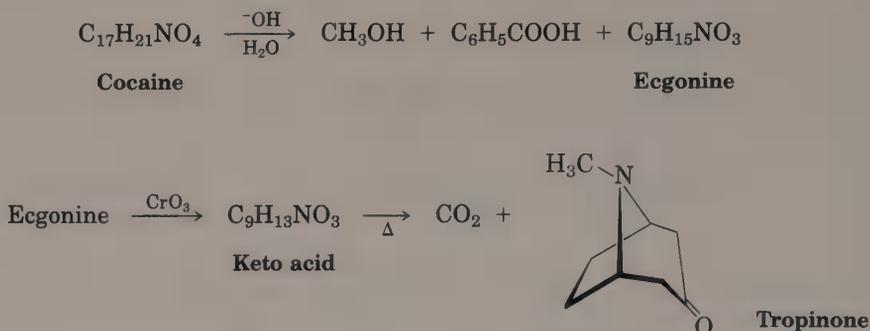


22.39 Although 2-substituted 2-cyclopentenones are in a base-catalyzed equilibrium with their 5-substituted 2-cyclopentenone isomers (Problem 22.38), the analogous isomerization is not observed for 2-substituted 2-cyclohexenones. Explain.



22.40 At least as far back as the sixteenth century, the Incas chewed the leaves of the coca bush, *Erythroxylon coca*, to combat fatigue. Chemical studies of *Erythroxylon coca* by Friedrich Wöhler in 1862 resulted in the discovery of *cocaine* as the active component. It was soon found that basic hydrolysis of cocaine led to methanol, benzoic

acid, and another compound called *ecgonine*. Chromium trioxide oxidation of *ecgonine* led to a keto acid that readily lost  $\text{CO}_2$  on heating, giving *tropinone*.

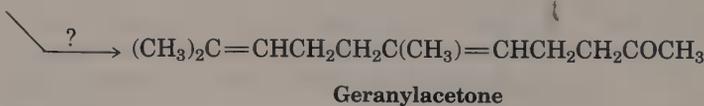
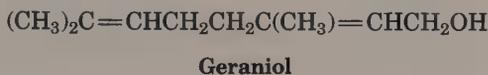
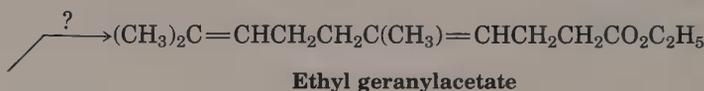


- What is a likely structure for the keto acid?
- What is a likely structure for *ecgonine*?
- What is a likely structure for *cocaine*?
- Formulate the reactions involved.

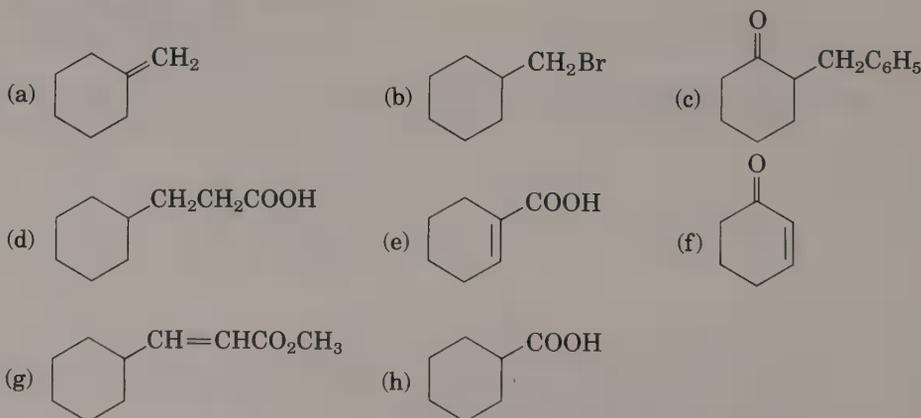
**22.41** Which of these substances would give a positive haloform reaction?

- $\text{CH}_3\text{COCH}_3$
- Acetophenone
- $\text{CH}_3\text{CH}_2\text{CHO}$
- $\text{CH}_3\text{COOH}$
- $\text{CH}_3\text{C}\equiv\text{N}$

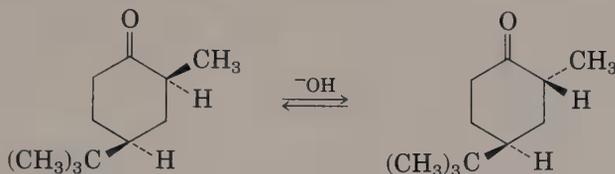
**22.42** Show how you might convert geraniol into either ethyl geranylacetate or geranylacetone.



**22.43** How would you synthesize the following compounds from cyclohexanone? More than one step may be required.

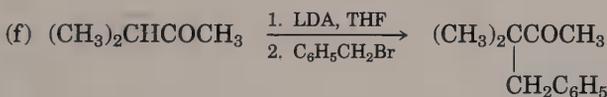
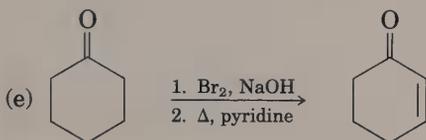
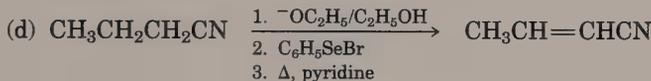
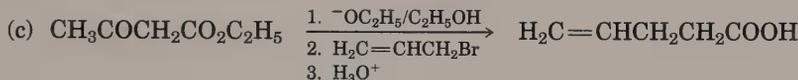
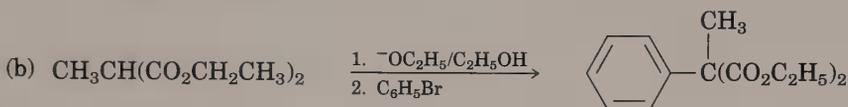
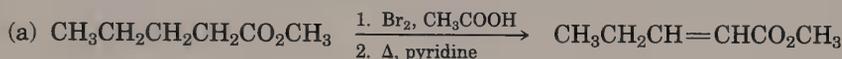


- 22.44 How can you account for the fact that *cis*- and *trans*-4-*tert*-butyl-2-methylcyclohexanone are interconverted by base treatment?

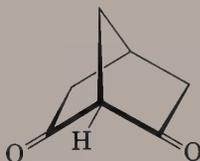


Which of the two isomers do you think is more stable, and why? Use molecular models to help formulate your answer.

- 22.45 The following synthetic routes are incorrect. What is wrong with each?

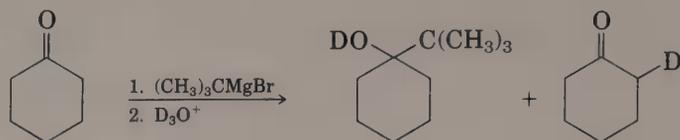


- 22.46 Unlike most  $\beta$ -diketones such as 2,4-pentanedione, the  $\beta$ -diketone shown has no detectable enol content and is about as acidic as acetone. Explain this behavior. Molecular models should be helpful.

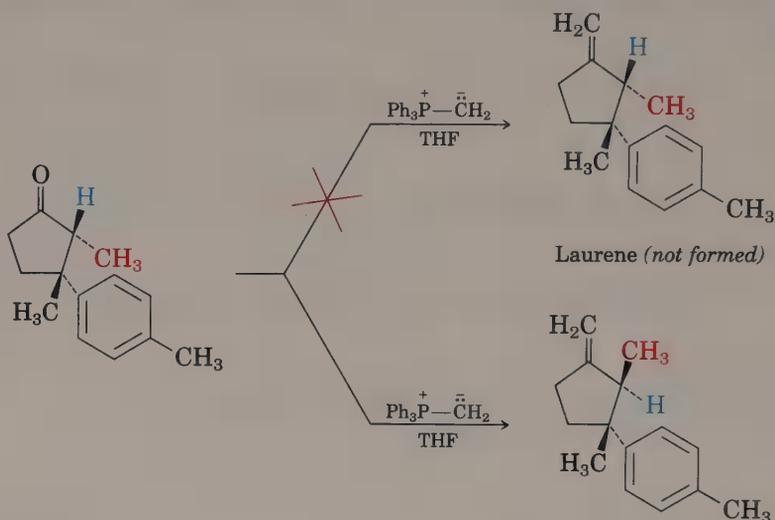


- 22.47 Methylmagnesium bromide adds to cyclohexanone to give the expected tertiary alcohol product in high yield. *tert*-Butylmagnesium bromide, however, gives only about a 1% yield of the addition product, along with 99% recovered starting material.

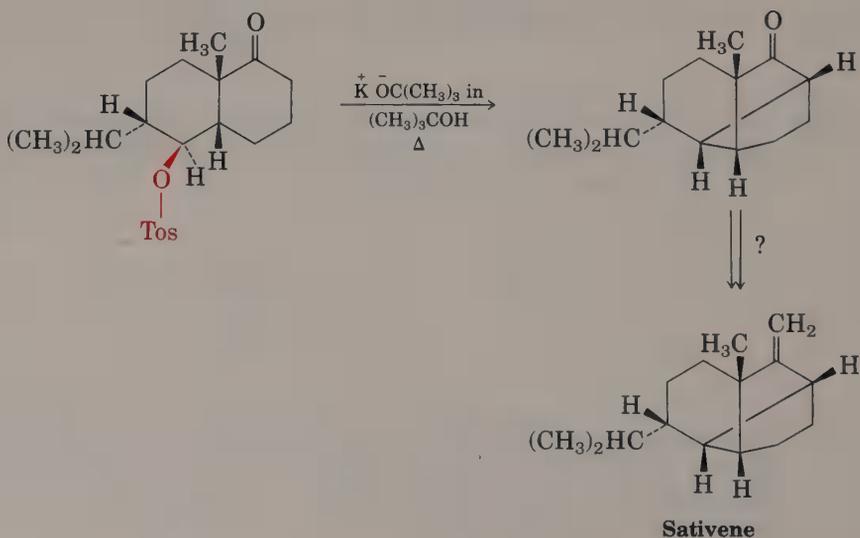
Furthermore, if  $D_3O^+$  is added to the reaction mixture after a suitable period, one deuterium atom is incorporated into the recovered cyclohexanone. Explain.



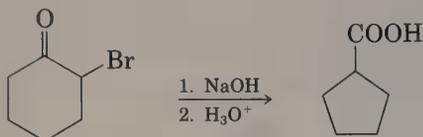
- 22.48 The final step in an attempted synthesis of laurene, a hydrocarbon isolated from the alga *Laurencia glandulifera*, involved the Wittig reaction shown. The product obtained, however, was not laurene but an isomer. Propose a mechanism to account for this unexpected result.



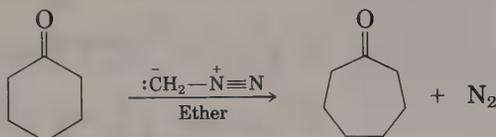
- 22.49 The key step in a reported laboratory synthesis of sativene, a hydrocarbon isolated from the mold *Helminthosporium sativum*, is shown. What kind of reaction is occurring? How would you complete the synthesis?

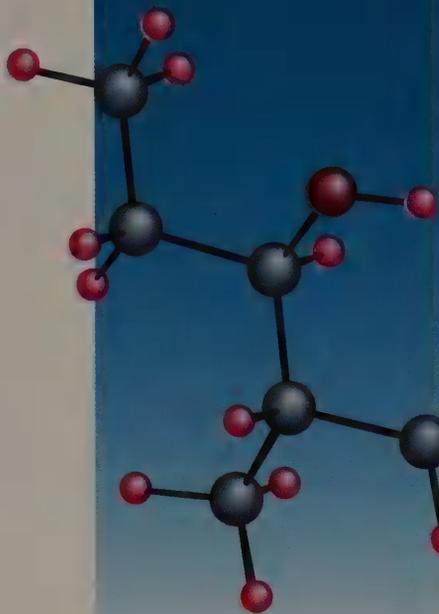


- 22.50** The Favorskii reaction involves treatment of an  $\alpha$ -bromo ketone with base to yield a ring-contracted product. For example, reaction of 2-bromocyclohexanone with aqueous NaOH yields cyclopentanecarboxylic acid. Propose a mechanism to account for this reaction.



- 22.51** Treatment of a cyclic ketone with diazomethane is a method for accomplishing a *ring-expansion reaction*. For example, treatment of cyclohexanone with diazomethane yields cycloheptanone. Propose a mechanism to account for this reaction.

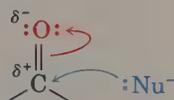




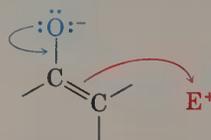
# Chapter 23

## Carbonyl Condensation Reactions

We've seen three general kinds of carbonyl-group reactions in the past four chapters and have studied two general kinds of behavior. In nucleophilic addition and nucleophilic acyl substitution reactions, the carbonyl group behaves as an electrophile by accepting electrons from an attacking nucleophile. In alpha-substitution reactions, however, the carbonyl compound behaves as a nucleophile when it is converted into its enolate ion or enol tautomer. The **carbonyl condensation reactions** that we'll study in the present chapter involve *both* types of reactivity.



Electrophilic carbonyl group  
is attacked by nucleophiles

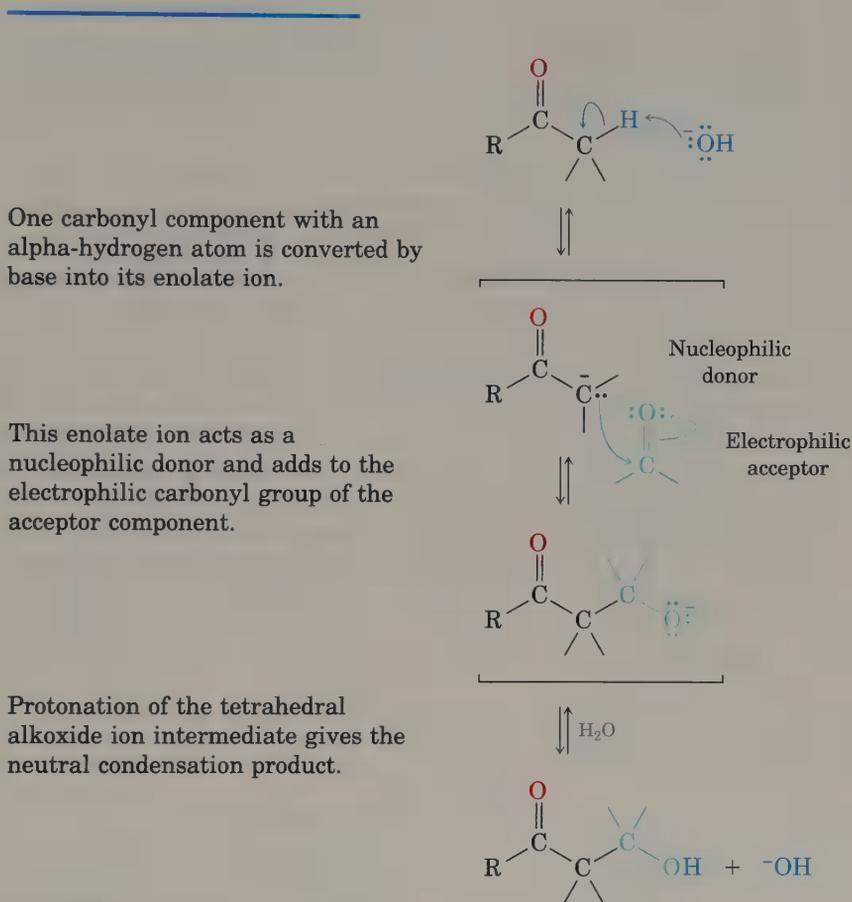


Nucleophilic enolate ion  
attacks electrophiles

### 23.1 General Mechanism of Carbonyl Condensation Reactions

Carbonyl condensation reactions take place between two carbonyl components and involve a *combination* of nucleophilic addition and alpha-substitution steps. One component acts as a nucleophilic electron donor in undergoing a nucleophilic addition, while the other component acts as an

electrophilic electron acceptor in undergoing an  $\alpha$ -substitution. The general mechanism of a carbonyl condensation reaction is shown in Figure 23.1.



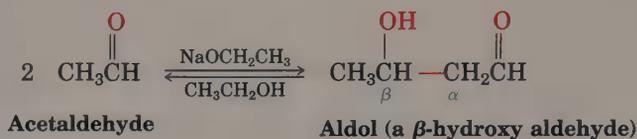
**Figure 23.1** The general mechanism of a carbonyl condensation reaction. One component (the donor) acts as a nucleophile, while the other component (the acceptor) acts as an electrophile.

Viewed from the side of the donor component, a carbonyl condensation reaction is simply an  $\alpha$ -substitution process. Viewed from the side of the acceptor component, a carbonyl condensation reaction is a nucleophilic addition process. Regardless of how the reaction is viewed, carbonyl condensations are among the most useful reactions in organic chemistry.

All manner of carbonyl compounds, including aldehydes, ketones, esters, amides, acid anhydrides, thiol esters, and nitriles, enter into condensation reactions. Nature uses carbonyl condensation reactions as key steps in the biosynthesis of many naturally occurring compounds, and chemists use the same reactions in the laboratory.

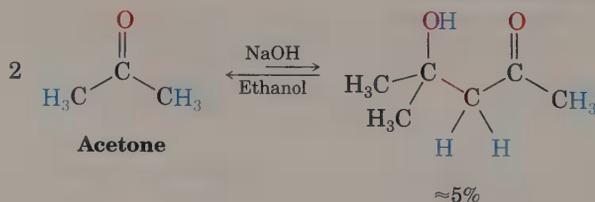
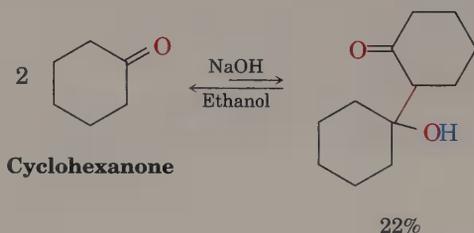
## 23.2 Condensations of Aldehydes and Ketones: The Aldol Reaction

When acetaldehyde is treated in a hydroxylic solvent with a basic catalyst such as sodium ethoxide or sodium hydroxide, a rapid and reversible condensation reaction occurs. The product is the  $\beta$ -hydroxy aldehyde known commonly as *aldol* (aldehyde + alcohol).

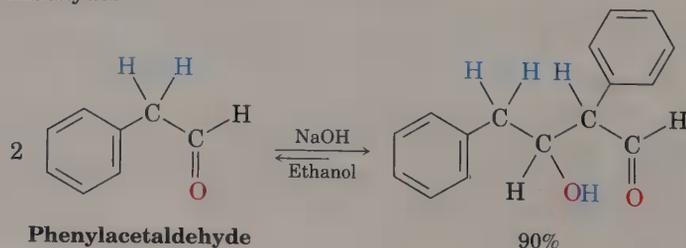


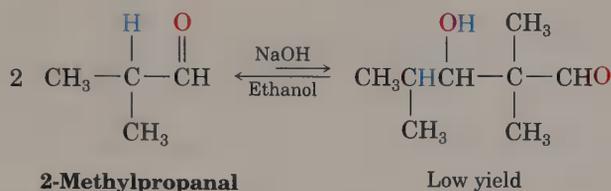
Called the **aldol reaction**, base-catalyzed dimerization is a general reaction for all ketones and aldehydes with  $\alpha$ -hydrogen atoms. If the ketone or aldehyde doesn't have an  $\alpha$ -hydrogen atom, however, aldol condensation can't occur. As the following examples indicate, the aldol equilibrium generally favors condensation product in the case of monosubstituted acetaldehydes ( $\text{RCH}_2\text{CHO}$ ), but favors starting material for disubstituted acetaldehydes ( $\text{R}_2\text{CHCHO}$ ) and for most ketones. Steric factors are probably responsible for these trends, since increased substitution near the reaction site greatly increases steric congestion in the aldol product.

### Ketones



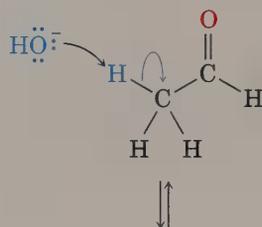
### Aldehydes



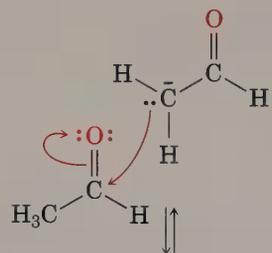


Aldol reactions are typical carbonyl condensations. They occur by nucleophilic addition of the enolate ion of the donor molecule to the carbonyl group of the acceptor molecule, yielding a tetrahedral intermediate that is protonated to give the alcohol as final product. The reverse process occurs in exactly the opposite manner: Base abstracts the hydroxyl proton to yield a  $\beta$ -keto alkoxide ion, which fragments to give one molecule of enolate ion and one molecule of neutral carbonyl compound (Figure 23.2).

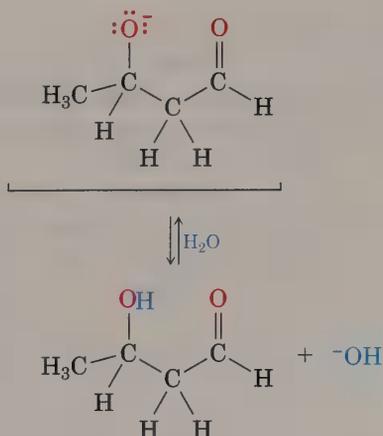
Base removes an acidic alpha hydrogen from one aldehyde molecule, yielding a resonance-stabilized enolate ion.



The enolate ion attacks a second aldehyde molecule in a nucleophilic addition reaction to give a tetrahedral alkoxide ion intermediate.



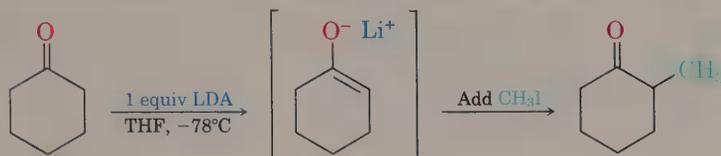
Protonation of the alkoxide ion intermediate yields neutral aldol product and regenerates the base catalyst.



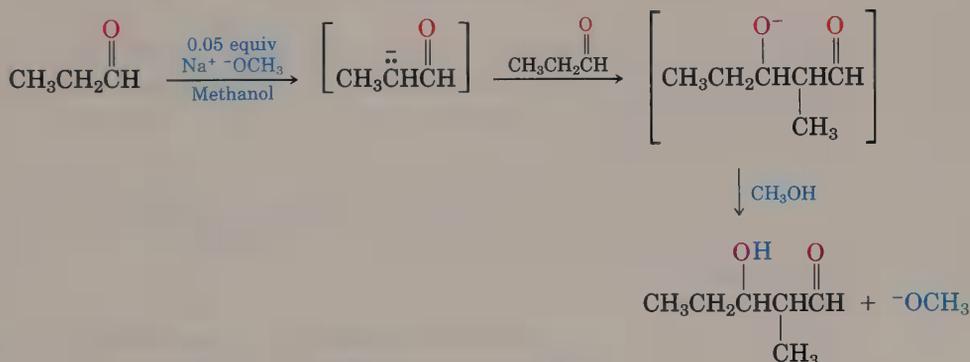
**Figure 23.2** Mechanism of the aldol reaction.



no condensation reaction could take place. We would then immediately add an alkyl halide to complete the alkylation reaction.

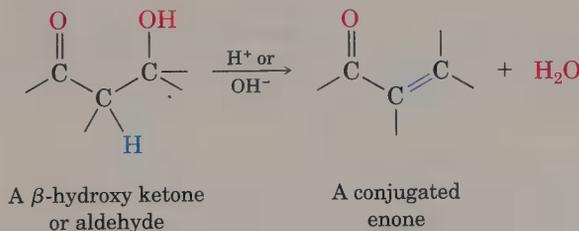


On the other hand, we might want to carry out a carbonyl condensation reaction. Since we need to generate only a small amount of the enolate ion in the presence of unreacted carbonyl compound, though, the aldol reaction requires only a *catalytic* amount of base, rather than a full equivalent. Once a condensation has occurred, the basic catalyst is regenerated. To carry out an aldol reaction on propanal, for example, we might dissolve the aldehyde in methanol, add 0.05 equivalent of sodium methoxide, and then warm the mixture. A high yield of aldol product would result.



## 23.4 Dehydration of Aldol Products: Synthesis of Enones

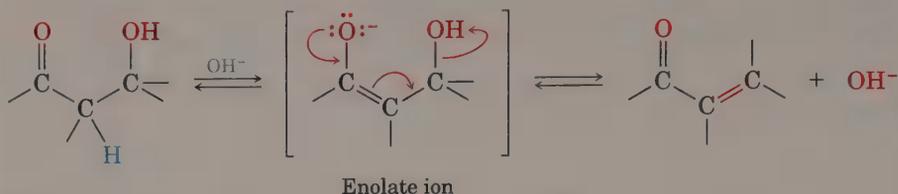
The  $\beta$ -hydroxy ketones and  $\beta$ -hydroxy aldehydes formed in aldol reactions can be easily dehydrated to yield conjugated enones. In fact, it's this loss of water that gives the aldol *condensation* its name, since water condenses out of the reaction when the enone product forms.



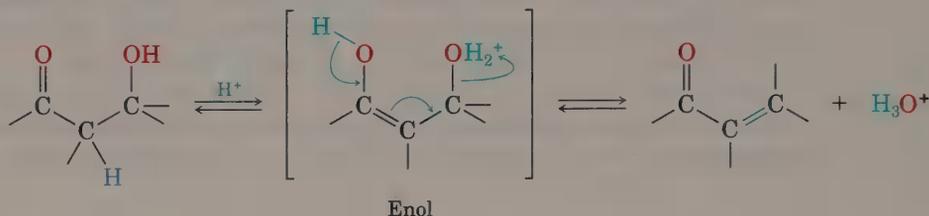
Most alcohols are resistant to dehydration by dilute acid or base, and special reagents like  $\text{POCl}_3$  must therefore be used (Section 17.8). Hydroxyl groups beta to a carbonyl group are special, however, because of the nearby carbonyl group. Under *basic* conditions, an acidic alpha hydrogen is ab-

stracted, yielding an enolate ion from which the hydroxide ion leaving group is expelled. Under *acidic* conditions, an enol is formed, the hydroxyl group is protonated, and water is expelled.

Base catalyzed

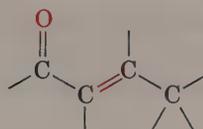


Acid catalyzed

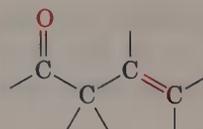


The conditions needed for aldol dehydration are often only a bit more vigorous (slightly higher temperature, for example) than the conditions needed for the aldol dimerization itself. As a result, conjugated enones are often obtained directly from aldol reactions, and the intermediate  $\beta$ -hydroxy carbonyl compounds aren't usually isolated.

Conjugated enones are more stable than nonconjugated enones for the same reasons that conjugated dienes are more stable than nonconjugated dienes (Section 14.2). Interaction between the pi electrons of the carbon-carbon double bond and the pi electrons of the carbonyl group leads to a molecular orbital description of conjugated enones that shows a partial delocalization of the pi electrons over all four atomic centers.

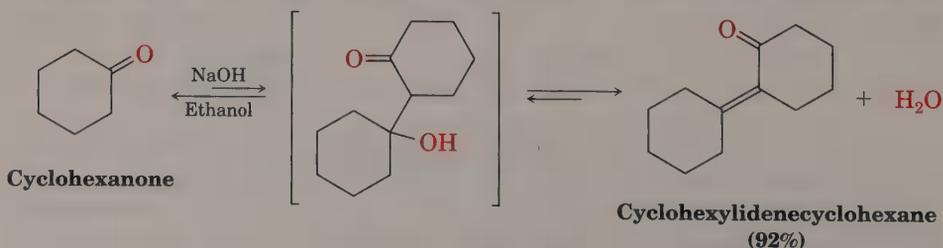


Conjugated enone  
(more stable)



Nonconjugated enone  
(less stable)

The real importance of aldol dehydration lies in the fact that removal of water from the reaction mixture acts to drive the equilibrium toward product formation. Even though the initial aldol step itself may be unfavorable (as it usually is for ketones), the subsequent dehydration step nevertheless allows most aldol condensations to be carried out in good yield.

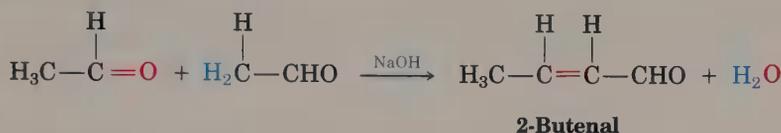


Cyclohexanone, for example, gives the dimer cyclohexylidenecyclohexanone in 92% yield even though the initial equilibrium is unfavorable.

**PRACTICE PROBLEM**.....

What is the structure of the enone obtained from aldol condensation of acetaldehyde?

**Solution** In the aldol reaction,  $\text{H}_2\text{O}$  is eliminated by removing two hydrogens from the acidic alpha position of one partner and the oxygen from the second partner:



**PROBLEM**.....

**23.4** What enone products would you expect from aldol condensation of the following compounds?

(a) Cyclopentanone

(b) Acetophenone

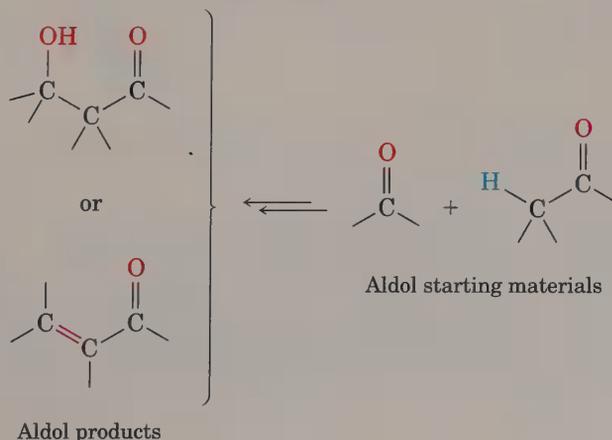
(c) 3-Methylbutanal

**PROBLEM**.....

**23.5** Aldol condensation of 3-methylcyclohexanone leads to a mixture of two products, not counting double-bond isomers. Draw them.

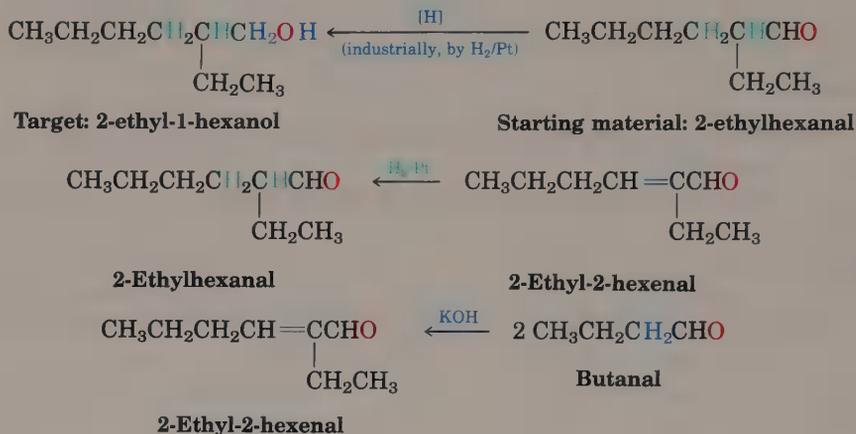
## 23.5 Recognizing Aldol Products

The aldol condensation reaction yields either  $\beta$ -hydroxy ketones/aldehydes or  $\alpha,\beta$ -unsaturated ketones/aldehydes, depending on the specific case and on the reaction conditions. By learning how to think *backward*, it's possible to predict when the aldol reaction might be useful in synthesis. Any time the target molecule contains either a  $\beta$ -hydroxy ketone/aldehyde or a conjugated enone functional group, it might come from an aldol reaction:



We can extend this kind of reasoning even further by considering that subsequent transformations might be carried out on the aldol products. For example, a saturated ketone might be prepared by catalytic hydrogenation of an enone product. A good example can be found in the industrial preparation of 2-ethyl-1-hexanol, an alcohol used in the synthesis of plasticizers.

Although 2-ethyl-1-hexanol bears little resemblance to an aldol product at first glance, it is prepared commercially from butanal by an aldol reaction. Working backward, we can reason that 2-ethyl-1-hexanol might come from 2-ethylhexanal by a reduction. 2-Ethylhexanal, in turn, might be prepared by catalytic reduction of 2-ethyl-2-hexenal, which is the aldol self-condensation product of butanal. The reactions that follow show the sequence in reverse order.



PROBLEM .....

- 23.6 Which of the following compounds are aldol self-condensation products? What is the ketone or aldehyde precursor of each?
- (a) 2,2,3-Trimethyl-3-hydroxybutanal      (b) 2-Methyl-2-hydroxypentanal  
 (c) 5-Ethyl-4-methyl-4-hepten-3-one

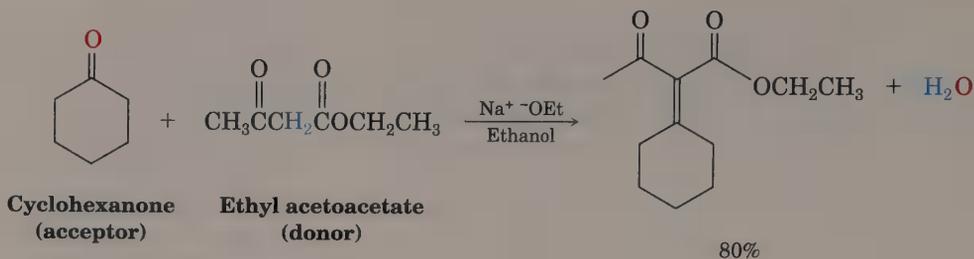
PROBLEM .....

- 23.7 1-Butanol is prepared commercially by a route that begins with the aldol reaction of acetaldehyde. Show the steps that are likely to be involved.
- .....

## 23.6 Mixed Aldol Reactions

Until now, we've considered only symmetrical aldol reactions, in which the two carbonyl components have been identical. What would happen, though, if a *mixed* aldol reaction were carried out between two different carbonyl partners?



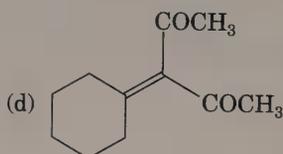
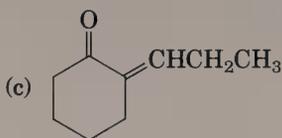
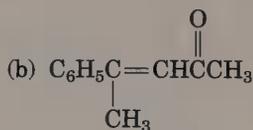
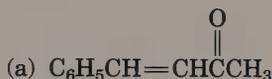


Ethyl acetoacetate is completely converted into its enolate ion in preference to enolate ion formation from other carbonyl partners such as cyclohexanone. Aldol condensation therefore occurs preferentially to give the mixed product.

The situation can be summarized by saying that a mixed aldol reaction between two different carbonyl partners leads to a mixture of products unless one of the partners either is an unusually good nucleophilic donor (such as acetoacetic ester) or else has no alpha protons and is a good electrophilic acceptor (such as benzaldehyde).

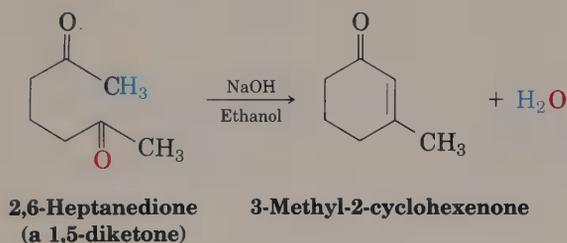
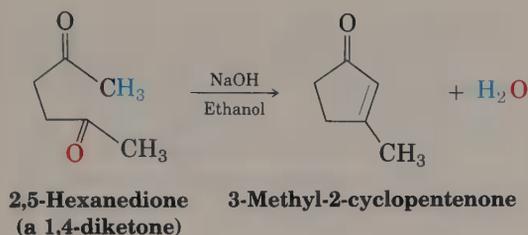
PROBLEM .....

- 23.8 Which of the following compounds can probably be prepared by a mixed aldol reaction?



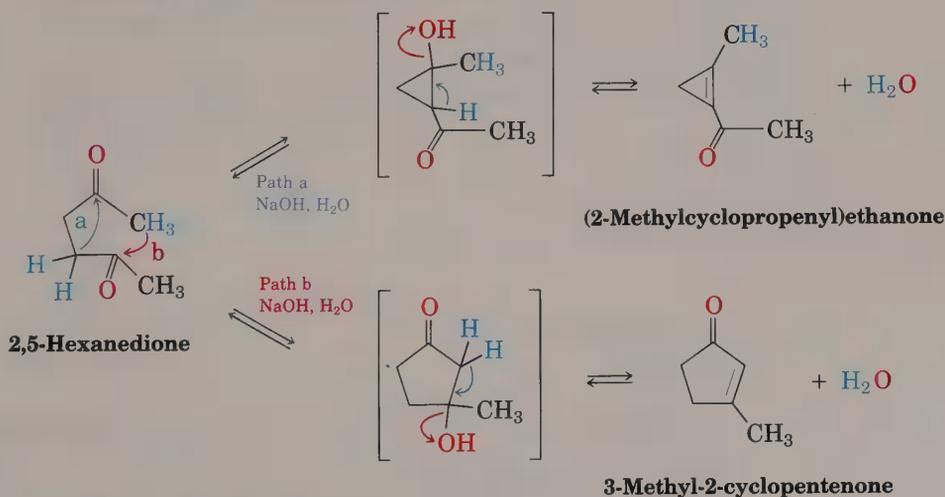
## 23.7 Intramolecular Aldol Reactions

The aldol reactions we've seen up to this point have all been *intermolecular*. That is, they have taken place between two different molecules. When certain *dicarbonyl* compounds are treated with base, however, *intramolecular* aldol reactions can occur, leading to the formation of cyclic products. For example, base treatment of a 1,4-diketone like 2,5-hexanedione yields a cyclopentenone product, and similar base treatment of a 1,5-diketone like 2,6-heptanedione yields a cyclohexenone product.



The mechanism of these intramolecular aldol reactions is similar to the mechanism of the corresponding intermolecular reactions. The only difference is that both the nucleophilic carbonyl anion donor and the electrophilic carbonyl acceptor are in the same molecule.

In principle, intramolecular aldol reactions can lead to a mixture of products, depending on which enolate ion is formed. For example, 2,5-hexanedione might yield either the five-membered-ring product 3-methyl-2-cyclopentenone or the three-membered-ring product (2-methylcyclopropenyl)ethanone (Figure 23.3). In practice, though, only the cyclopentenone is formed.



**Figure 23.3** Intramolecular aldol reaction of 2,5-hexanedione yields 3-methyl-2-cyclopentenone rather than the alternative acetylcyclopropene.

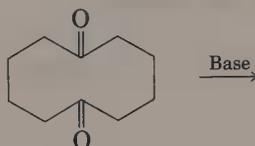
The observed product selectivity in intramolecular aldol reactions is due to the fact that all steps in the mechanism are reversible and that an equilibrium is reached. Thus, the relatively strain-free cyclopentenone product is considerably more stable than the highly strained cyclopropene alternative. For similar reasons, intramolecular aldol reactions of 1,5-diketones lead only to cyclohexenone products rather than to strained cyclobutenes.

PROBLEM .....

- 23.9 Why do you suppose that 1,3-diketones do not undergo internal aldol condensation to yield cyclobutenones?

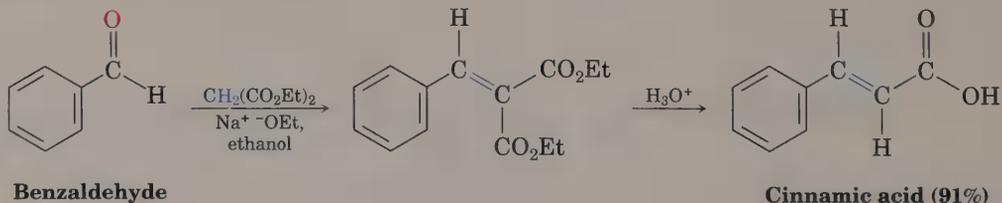
PROBLEM .....

- 23.10 What product would you expect to obtain from base treatment of 1,6-cyclodecane-1,6-dione?



## 23.8 Reactions Similar to the Aldol Condensation

The aldol reaction is usually defined as the condensation of aldehydes and ketones. In a more general sense, however, there are many similar reactions in which a compound with acidic alpha hydrogens condenses with a second carbonyl component. For example, diethyl malonate can condense with aldehydes and unhindered ketones to yield  $\alpha,\beta$ -unsaturated diesters—the **Knoevenagel<sup>1</sup> reaction**. Like other malonates, these diesters can then be decarboxylated (Section 22.9).



Many other possibilities exist, including condensations with nitriles, anhydrides, and nitro compounds as the acidic donor components. The many

<sup>1</sup>Emil Knoevenagel (1865–1921); b. Linden/Hannover, Germany; Ph.D. Göttingen; professor, University of Heidelberg.

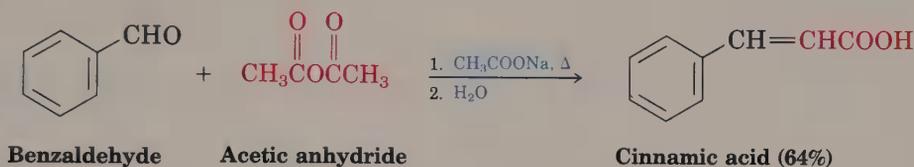
possibilities lead to an extremely rich and varied chemistry, but all are simply variations of the generalized aldol mechanism in which a nucleophilic carbon donor adds to an electrophilic carbonyl acceptor.

## PROBLEM

- 23.11 Show the mechanism of the Knoevenagel reaction of diethyl malonate and benzaldehyde.

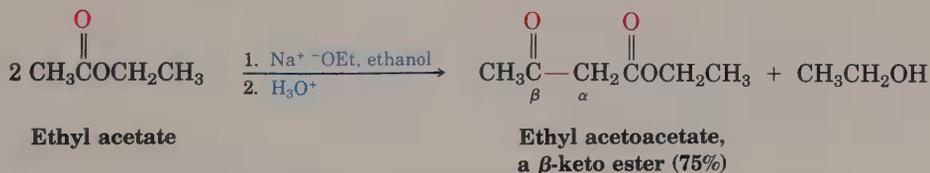
## PROBLEM

- 23.12 In the *Perkin reaction*, acetic anhydride condenses with an aromatic aldehyde to yield a cinnamic acid. The reaction takes place by a mixed carbonyl condensation of the anhydride with the aldehyde to yield an  $\alpha,\beta$ -unsaturated intermediate that undergoes hydrolysis to yield the cinnamic acid. What is the structure of the unsaturated intermediate?



## 23.9 The Claisen Condensation Reaction

Esters, like aldehydes and ketones, are weakly acidic. When an ester with an alpha hydrogen is treated with one equivalent of a base such as sodium ethoxide, a reversible condensation reaction occurs to yield a  $\beta$ -keto ester product. For example, ethyl acetate yields ethyl acetoacetate on treatment with base. This reaction between two ester components is known as the **Claisen<sup>2</sup> condensation reaction**.



The mechanism of the Claisen reaction is similar to that of the aldol reaction. As shown in Figure 23.4, the Claisen condensation involves the nucleophilic addition of an ester enolate-ion donor to the carbonyl group of a second ester molecule. From the point of view of the donor component, the

<sup>2</sup>Ludwig Claisen (1851–1930); b. Cologne; Ph.D. Bonn (Kekulé); professor, University of Bonn, Owens College (Manchester), universities of Munich, Aachen, Kiel, and Berlin; Godesberg (private laboratory).

reaction is simply an alpha substitution; from the point of view of the acceptor component, the reaction is a nucleophilic acyl substitution.

The only difference between an aldol condensation and a Claisen condensation involves the fate of the initially formed tetrahedral intermediate. The tetrahedral intermediate in the aldol reaction is protonated to give an alcohol product—exactly the behavior previously seen for ketones and aldehydes (Section 19.6). The tetrahedral intermediate in the Claisen reaction

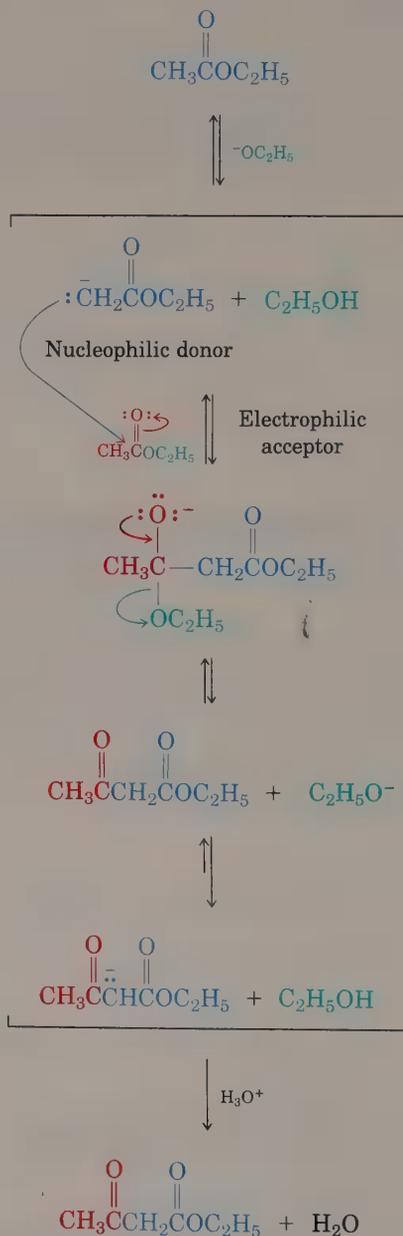
Ethoxide base abstracts an acidic alpha hydrogen atom from an ester molecule, yielding an ester enolate ion.

This ion does a nucleophilic addition to a second ester molecule, giving a tetrahedral intermediate.

The tetrahedral intermediate is not stable. It expels ethoxide ion to yield the new carbonyl compound, ethyl acetoacetate.

But ethoxide ion is a base. It therefore converts the  $\beta$ -keto ester product into its enolate, thus shifting the equilibrium and driving the reaction to completion.

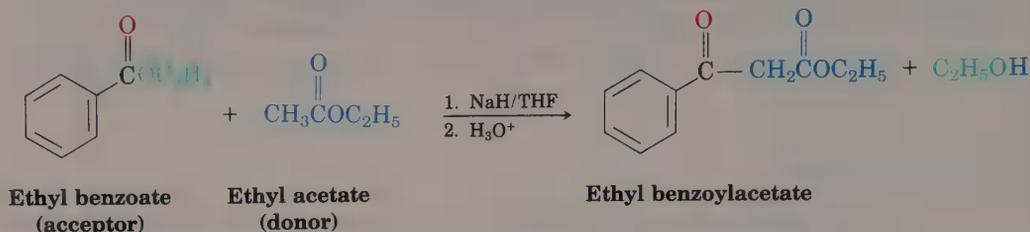
Protonation by addition of acid yields the final product.



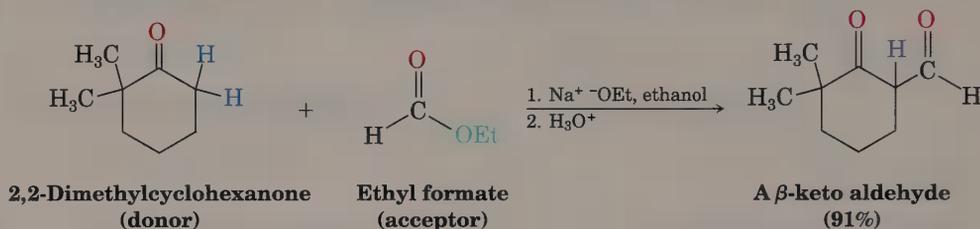
**Figure 23.4** Mechanism of the Claisen condensation reaction.



the two ester components has no alpha hydrogens and thus can't form an enolate ion. For example, ethyl benzoate and ethyl formate can't form enolate ions and thus can't serve as donors. They can, however, act as the electrophilic acceptor components in reactions with other ester anions to give good yields of mixed  $\beta$ -keto ester products.



Mixed Claisen-like reactions can also be carried out between esters and ketones. The result is an excellent synthesis of  $\beta$ -diketones. The reaction works best when the ester component has no alpha hydrogens and thus can't act as the nucleophilic donor. For example, ethyl formate gives particularly high yields in mixed Claisen condensations with ketones.



#### PROBLEM.....

- 23.15 Would you expect diethyl oxalate,  $(\text{CO}_2\text{Et})_2$ , to give good yields in mixed Claisen reactions? Explain your answer. What product would you expect to obtain from mixed Claisen reaction of ethyl acetate with diethyl oxalate?

#### PROBLEM.....

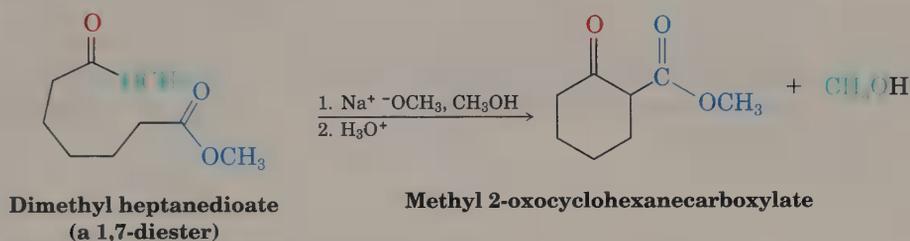
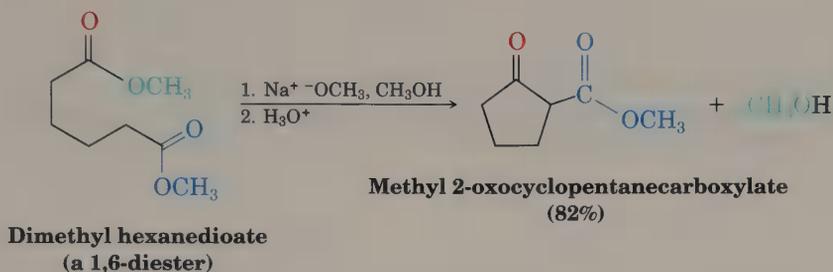
- 23.16 What product would you expect from a mixed Claisen-like reaction of 2,2-dimethylcyclohexanone with diethyl oxalate (Problem 23.15)?

## 23.11 Intramolecular Claisen Condensations: The Dieckmann Cyclization

Intramolecular Claisen condensations can be carried out with diesters, just as intramolecular aldol condensations can be carried out with diketones (Section 23.7). Called the **Dieckmann<sup>3</sup> cyclization**, the reaction works best

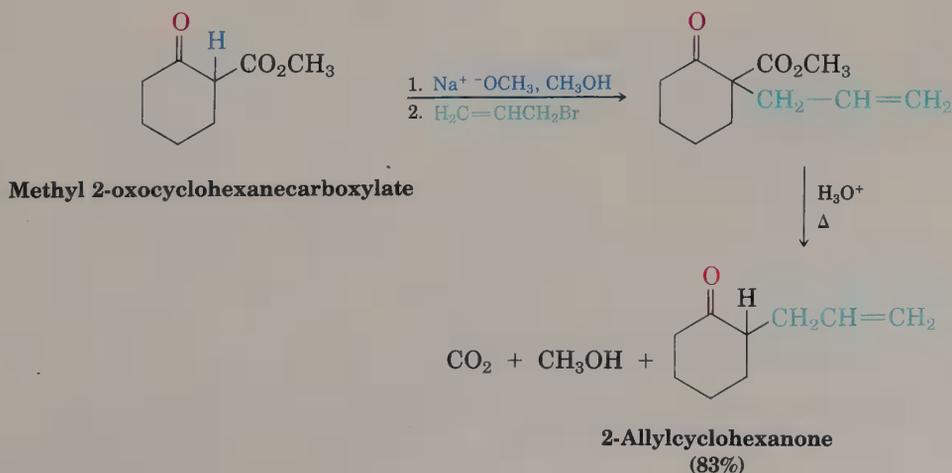
<sup>3</sup>Walter Dieckmann (1869–1925); b. Hamburg, Germany; Ph.D. Munich (Bamberger); professor, University of Munich.

on 1,6-diesters and 1,7-diesters. Five-membered-ring cyclic  $\beta$ -keto esters result from Dieckmann cyclization of 1,6-diesters, and six-membered-ring  $\beta$ -keto esters result from cyclization of 1,7-diesters.

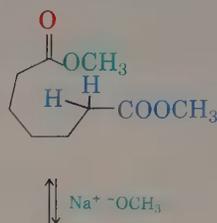


The mechanism of the Dieckmann cyclization, shown in Figure 23.5 on page 898, is analogous to the Claisen reaction. One of the two ester groups is converted into an enolate ion, which then carries out a nucleophilic attack on the second ester group at the other end of the molecule. A cyclic  $\beta$ -keto ester product results.

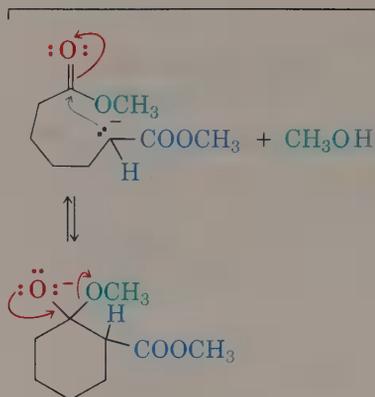
The products of the Dieckmann cyclization are cyclic  $\beta$ -keto esters that can be further alkylated and decarboxylated by a series of reactions analogous to the acetoacetic ester synthesis (Section 22.9). For example, alkylation and subsequent decarboxylation of methyl 2-oxocyclohexanecarboxylate yields a 2-alkylcyclohexanone. The overall sequence of Dieckmann cyclization,  $\beta$ -keto ester alkylation, and decarboxylation is an excellent method for preparing 2-substituted cyclohexanones and cyclopentanones.



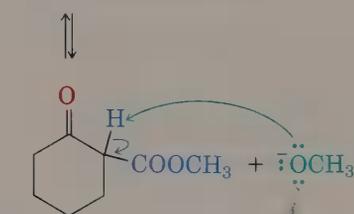
Base abstracts an acidic alpha proton from the carbon atom next to one of the ester groups, yielding an enolate ion.



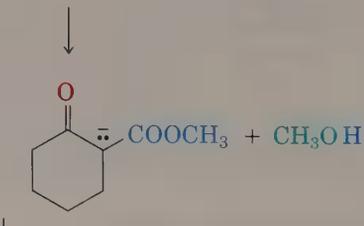
Intramolecular nucleophilic addition of the ester enolate ion to the carbonyl group of the second ester group at the other end of the chain then gives a cyclic tetrahedral intermediate.



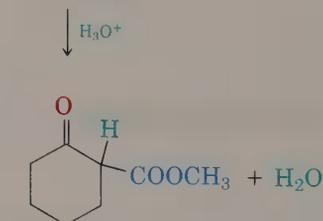
Loss of alkoxide ion from the tetrahedral intermediate forms a cyclic  $\beta$ -keto ester.



Deprotonation of the acidic  $\beta$ -keto ester gives an enolate ion . . .



. . . which is protonated by addition of aqueous acid at the end of the reaction to generate the neutral  $\beta$ -keto ester product.



**Figure 23.5** Mechanism of the Dieckmann cyclization of a diester to yield a cyclic  $\beta$ -keto ester product.

PROBLEM .....

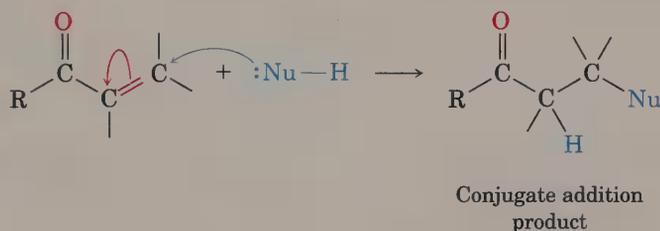
- 23.17 What product would you expect on treatment of diethyl 4-methylheptanedioate with sodium ethoxide followed by acidification?

PROBLEM .....

- 23.18 How can you account for the fact that Dieckmann cyclization of diethyl 3-methylheptanedioate gives a mixture of two  $\beta$ -keto ester products? What are their structures, and why is a mixture formed?

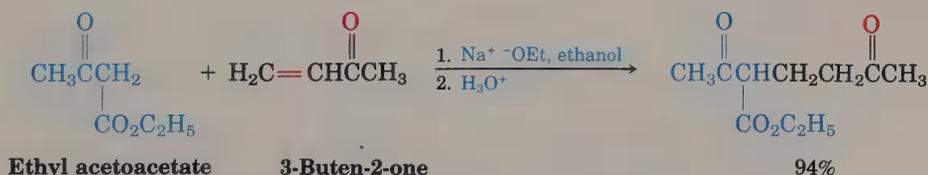
## 23.12 The Michael Reaction

We saw earlier (Section 19.18) that nucleophiles such as amines and cyanide ion can react with  $\alpha,\beta$ -unsaturated ketones and aldehydes to give the conjugate addition product, rather than the direct addition product:



Exactly the same kind of conjugate addition can occur when a nucleophilic enolate ion reacts with an  $\alpha,\beta$ -unsaturated carbonyl compound—a process known as the **Michael<sup>4</sup> reaction**.

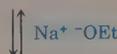
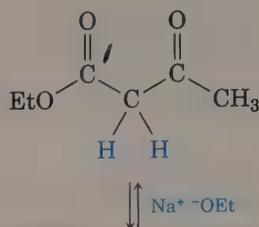
The best Michael reactions take place when particularly stable enolate ions such as those derived from  $\beta$ -keto esters or  $\beta$ -diesters (malonic esters) add to unhindered  $\alpha,\beta$ -unsaturated ketones. For example, ethyl acetoacetate reacts with 3-buten-2-one in 94% yield in the presence of sodium ethoxide catalyst to yield the conjugate addition product:



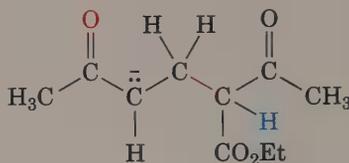
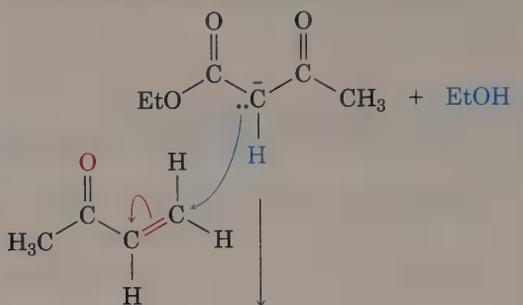
Michael reactions take place by addition of a nucleophilic enolate ion donor to the beta carbon of an  $\alpha,\beta$ -unsaturated carbonyl acceptor, according to the mechanism shown in Figure 23.6.

<sup>4</sup>Arthur Michael (1853–1942); b. Buffalo, New York; studied Heidelberg, Berlin, École de Médecine, Paris; professor, Tufts University (1882–1889 and 1894–1907), Harvard University (1912–1936).

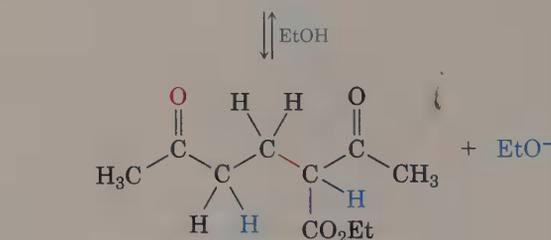
The base catalyst removes an acidic alpha proton from the starting  $\beta$ -keto ester to generate a stabilized enolate ion nucleophile.



The nucleophile adds to the  $\alpha,\beta$ -unsaturated ketone electrophile in a Michael reaction to generate a new enolate as product.



The enolate product abstracts an acidic proton, either from solvent or from starting keto ester, to yield the final addition product.



**Figure 23.6** Mechanism of the Michael reaction between a  $\beta$ -keto ester and an  $\alpha,\beta$ -unsaturated ketone.

The Michael reaction is general for a wide variety of  $\alpha,\beta$ -unsaturated carbonyl compounds, not just conjugated enones. Conjugated aldehydes, esters, nitriles, amides, and nitro compounds can all act as the electrophilic acceptor component in the Michael reaction (Table 23.1). Similarly, a variety of different donors can be used, including  $\beta$ -diketones,  $\beta$ -keto esters, malonic esters,  $\beta$ -keto nitriles, and nitro compounds.

**PROBLEM** .....

- 23.19** What products would you obtain from base-induced Michael reaction of 2,4-pentanedione with these  $\alpha,\beta$ -unsaturated acceptors?  
 (a) 2-Cyclohexenone      (b) Propenenitrile      (c) Methyl 2-butenoate

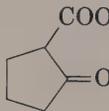
TABLE 23.1 Some Michael Acceptors and Michael Donors

<i>Michael acceptors</i>		<i>Michael donors</i>	
$\text{H}_2\text{C}=\text{CHCHO}$	Propenal	$\text{RCOCH}_2\text{COR}'$	$\beta$ -Diketone
$\text{H}_2\text{C}=\text{CHCO}_2\text{CH}_3$	Methyl propenoate	$\text{RCOCH}_2\text{CO}_2\text{CH}_3$	$\beta$ -Keto ester
$\text{H}_2\text{C}=\text{CHC}\equiv\text{N}$	Propenenitrile	$\text{CH}_3\text{O}_2\text{CCH}_2\text{CO}_2\text{CH}_3$	Malonic ester
$\text{H}_2\text{C}=\text{CHCOCH}_3$	3-Buten-2-one	$\text{RCOCH}_2\text{C}\equiv\text{N}$	$\beta$ -Keto nitrile
$\text{H}_2\text{C}=\text{CHNO}_2$	Nitroethylene	$\text{RCH}_2\text{NO}_2$	Nitro compound
$\text{H}_2\text{C}=\text{CHCONH}_2$	Propenamide		

## PROBLEM .....

23.20 What products would you obtain from base-induced Michael reaction of 3-buten-2-one with these nucleophilic donors?

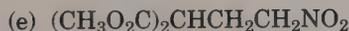
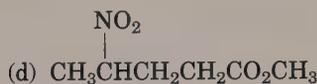
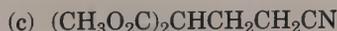
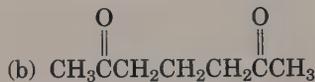
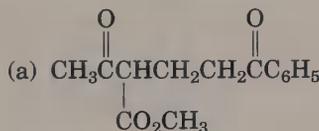
(a) Diethyl malonate

(b) 

(c) Nitromethane

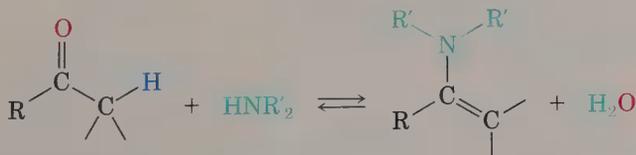
## PROBLEM .....

23.21 How might the following compounds be prepared using Michael reactions? Show the nucleophilic donor and the electrophilic acceptor in each case.

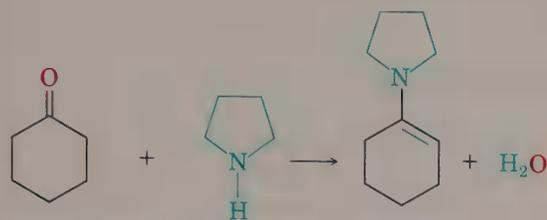


## 23.13 The Stork Enamine Reaction

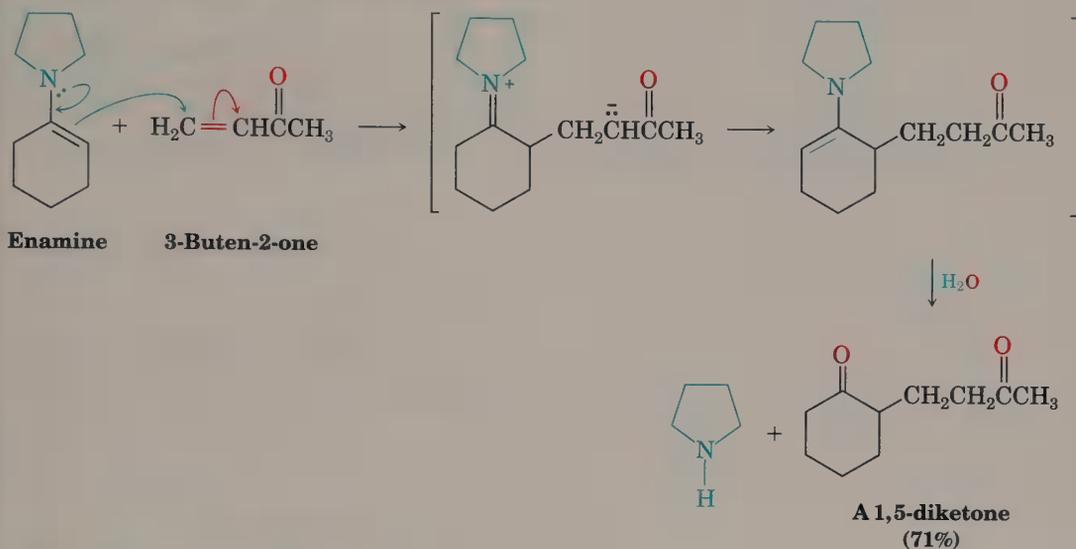
Other kinds of carbon nucleophiles besides enolate ions add to  $\alpha,\beta$ -unsaturated acceptors in the Michael reaction, greatly extending the usefulness and versatility of the process. Among the most important such nucleophiles are *enamines*. Recall from Section 19.12 that enamines are readily prepared by reaction between a ketone and a secondary amine.







Cyclohexanone    Pyrrolidine    An enamine



Enamine    3-Buten-2-one

A 1,5-diketone  
(71%)

**Figure 23.7** A Stork enamine reaction between cyclohexanone and 3-buten-2-one. Cyclohexanone is first converted into an enamine; the enamine then adds to the  $\alpha,\beta$ -unsaturated ketone in a Michael reaction; and the initial product is hydrolyzed to yield a 1,5-diketone.

## PROBLEM .....

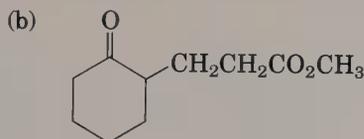
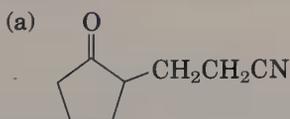
**23.23** What products would result from reaction of each enamine you prepared in Problem 23.22 with these  $\alpha,\beta$ -unsaturated acceptors? (Assume that the initial product is hydrolyzed.)

(a) Ethyl propenoate

(b) Propenal (acrolein)

## PROBLEM .....

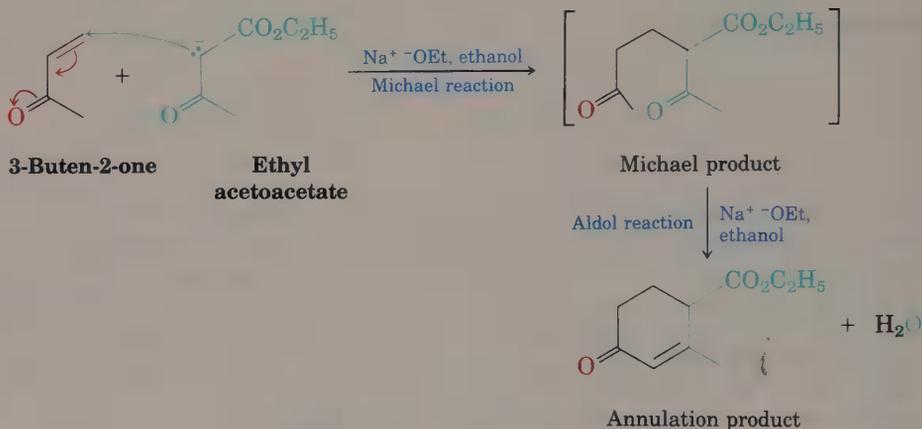
**23.24** Show how you might use an enamine reaction to prepare these compounds:



## 23.14 Carbonyl Condensation Reactions in Synthesis: The Robinson Annulation Reaction

Carbonyl condensation reactions are among the most valuable methods available for synthesizing complex molecules. By putting a few fundamental reactions together in the proper sequence, some remarkably useful transformations can be carried out. One such example is the **Robinson<sup>6</sup> annulation reaction**, used for the synthesis of polycyclic molecules. (An *annulation* reaction [from the Latin *annulus*, meaning “ring”] is one that builds a new ring onto a molecule.)

The Robinson annulation is a two-step process that combines a Michael reaction with an internal aldol reaction. It takes place between a Michael-type nucleophilic donor such as a  $\beta$ -keto ester or  $\beta$ -diketone and an  $\alpha,\beta$ -unsaturated ketone acceptor such as 3-buten-2-one. The product is a substituted 2-cyclohexenone:

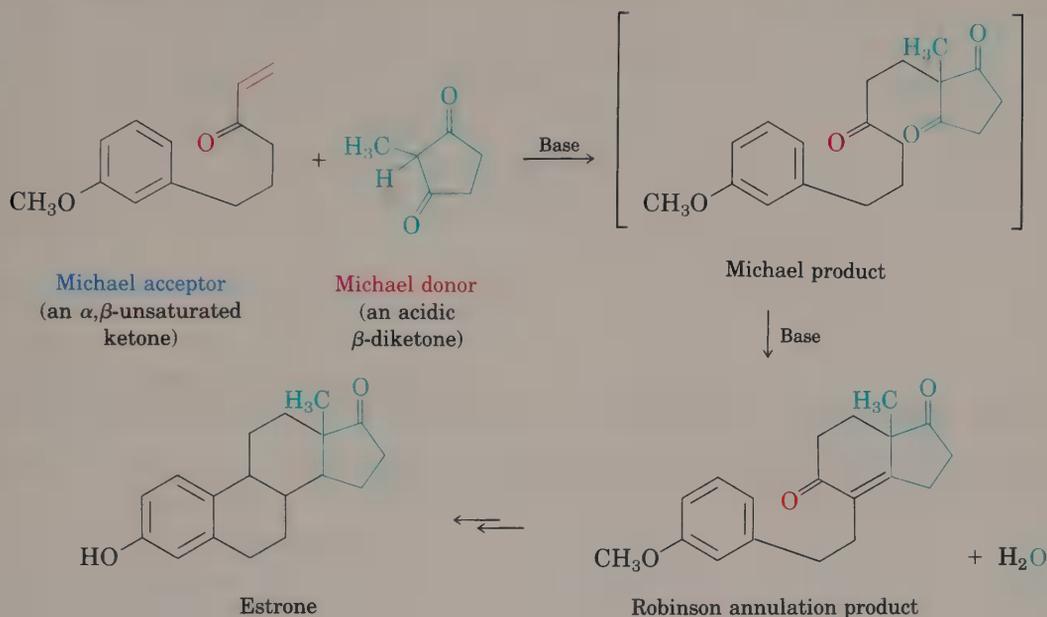


The first step of the Robinson annulation is simply a Michael reaction: a stabilized enolate ion from a  $\beta$ -diketone or  $\beta$ -keto ester effects a conjugate addition to an  $\alpha,\beta$ -unsaturated ketone, yielding a 1,5-diketone. But as we saw in Section 23.7, 1,5-diketones undergo intramolecular aldol condensation to yield cyclohexenones when treated with base. Thus, if base treatment of the initial product is continued at a higher temperature, the final product contains a six-membered ring and an annulation has been accomplished.

An excellent example of the practical importance of the Robinson annulation reaction occurs as a key step during the commercial synthesis of the female steroid hormone estrone (Figure 23.8).

In this example, 2-methyl-1,3-cyclopentanedione (a  $\beta$ -diketone) is used to generate the stabilized enolate ion required for Michael reaction, and an aryl-substituted  $\alpha,\beta$ -unsaturated ketone is used as the acceptor. Base-

<sup>6</sup>Sir Robert Robinson (1886–1975); b. Rufford/Chesterfield, England; D.Sc. Manchester (Perkin); professor, Liverpool, Manchester (1922–1928), University College, Oxford (1930–1955); Nobel Prize (1947).

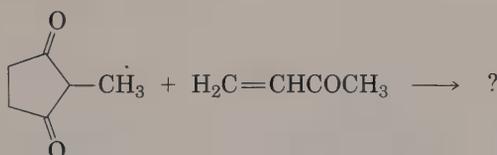


**Figure 23.8** A Robinson annulation reaction used in the commercial synthesis of the steroid hormone estrone.

catalyzed Michael reaction between the two yields an intermediate triketone, which immediately cyclizes in an internal aldol condensation to give a Robinson annulation product. Several further transformations are then required to complete the synthesis of estrone, but the Robinson annulation serves as the key step for assembling much of the molecule.

**PROBLEM** .....

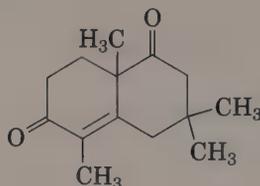
- 23.25** What product would you expect from a Robinson annulation reaction of 2-methyl-1,3-cyclopentanedione and 3-buten-2-one?



**PROBLEM** .....

- 23.26** How would you prepare the following compound using a Robinson annulation reaction between a  $\beta$ -diketone and an  $\alpha,\beta$ -unsaturated ketone acceptor? Draw the struc-

tures of both reactants you would use and the structure of the intermediate Michael addition product.



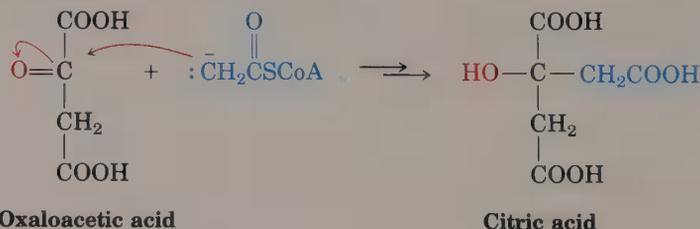
## 23.15 Biological Carbonyl Condensation Reactions

Carbonyl condensation reactions are used in nature for the biological synthesis of a great many different molecules. Fats, amino acids, steroid hormones, and many other kinds of compounds are synthesized by plants and animals using carbonyl condensation reactions as the key step.

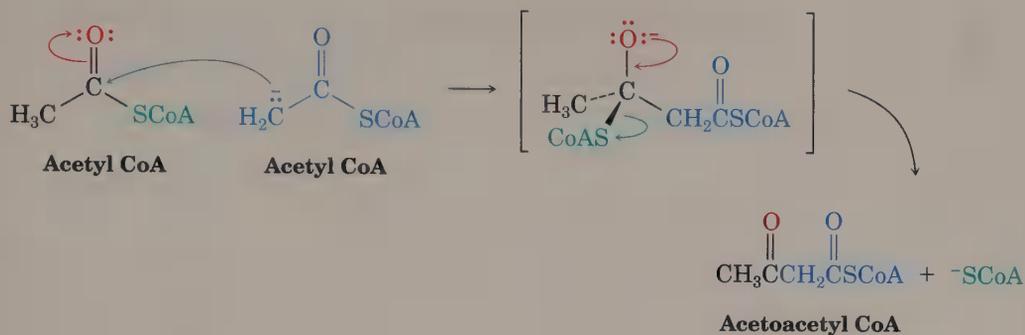
Nature uses the two-carbon acetate fragment of acetyl CoA as the major building block for synthesis. Acetyl CoA not only can serve as an electrophilic acceptor for attack of nucleophiles at the acyl carbon; it can also serve as a nucleophilic donor by loss of its acidic alpha proton to generate an enolate ion. The enolate ion of acetyl CoA can then add to another carbonyl group in a condensation reaction. For example, citric acid is biosynthesized by addition of acetyl CoA to the ketone carbonyl group of oxaloacetic acid (2-oxobutanedioic acid) in a kind of mixed aldol reaction, followed by hydrolysis of the thiol ester group.



Acetyl CoA,  
■ thiol ester

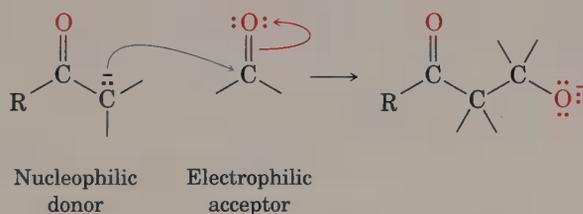


Acetyl CoA is also involved as a primary building block in the biosynthesis of steroids, fats, and other lipids, where the key step is a Claisen-like condensation reaction. We'll go into more of the details of this process in Section 28.4.

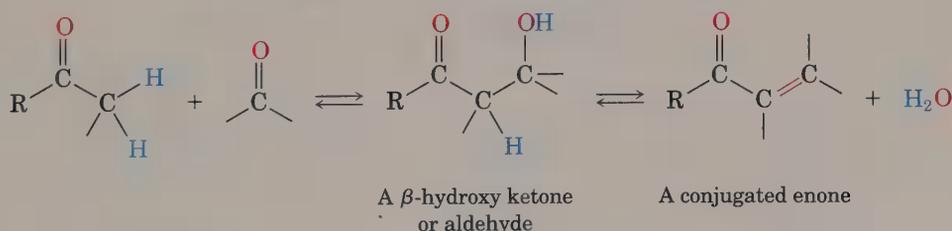


## 23.16 Summary and Key Words

A **carbonyl condensation reaction** takes place between two carbonyl components and involves a combination of nucleophilic addition and alpha-substitution steps. One carbonyl component (the donor) is converted by base into a nucleophilic enolate ion, which adds to the electrophilic carbonyl group of the second component (the acceptor). The donor molecule undergoes an alpha substitution, while the acceptor molecule undergoes a nucleophilic addition.

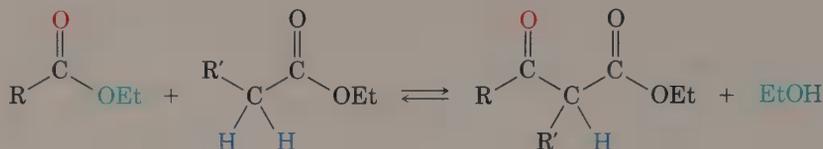


The **aldol reaction** is a carbonyl condensation that occurs between two ketone or aldehyde components. Aldol reactions are reversible, leading first to  $\beta$ -hydroxy ketones (or aldehydes) and then to  $\alpha,\beta$ -unsaturated ketones.



Mixed aldol condensations between two different ketones generally give a mixture of all four possible products. A mixed reaction can be successful, however, if one of the two components is an unusually good donor (as with ethyl acetoacetate) or if it can act only as an acceptor (as with formaldehyde and benzaldehyde). Internal aldol condensations of 1,4- and 1,5-diketones are also successful and provide an excellent method for preparing five- and six-membered enone rings.

The **Claisen reaction** is a carbonyl condensation that occurs between two ester components and leads to a  $\beta$ -keto ester product:



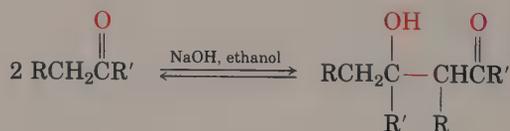
Mixed Claisen condensations between two different esters are successful only when one of the two components has no acidic alpha hydrogens (as with ethyl benzoate and ethyl formate) and thus can function only as the acceptor component. Internal Claisen condensations (**Dieckmann cyclization reactions**) provide excellent syntheses of five- and six-membered cyclic  $\beta$ -keto esters, starting from 1,6- and 1,7-diester.

The conjugate addition of carbon nucleophiles to an  $\alpha,\beta$ -unsaturated acceptor is known as the **Michael reaction**. The best Michael reactions take place between unusually acidic donors ( $\beta$ -keto esters or  $\beta$ -diketones) and unhindered  $\alpha,\beta$ -unsaturated acceptors. Enamines, prepared by reaction of a ketone with a disubstituted amine, are also excellent Michael donors.

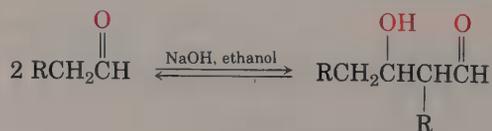
Carbonyl condensation reactions enjoy widespread use in synthesis. One example of their versatility is the **Robinson annulation reaction**, which leads to the formation of substituted cyclohexenone products. Treatment of a  $\beta$ -diketone or  $\beta$ -keto ester with an  $\alpha,\beta$ -unsaturated ketone leads first to a Michael addition, which is followed by internal aldol cyclization of the product. Condensation reactions are also used widely in nature for the biosynthesis of such molecules as fats and steroids.

## 23.17 Summary of Reactions

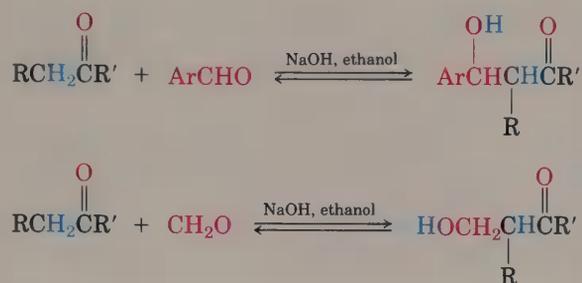
1. Aldol reaction: a condensation between two ketones, two aldehydes, or one ketone and one aldehyde
  - a. Ketones (Section 23.2)



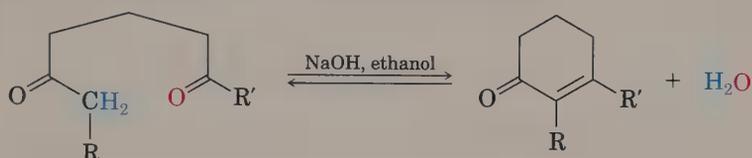
- b. Aldehydes (Section 23.2)



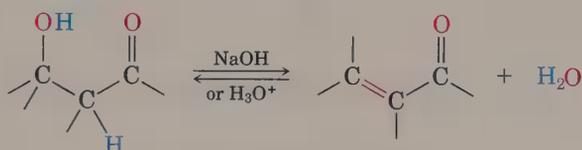
c. Mixed aldol reaction (Section 23.6)



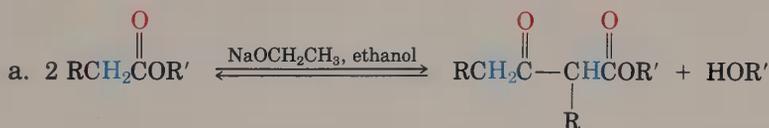
d. Intramolecular aldol reaction (Section 23.7)



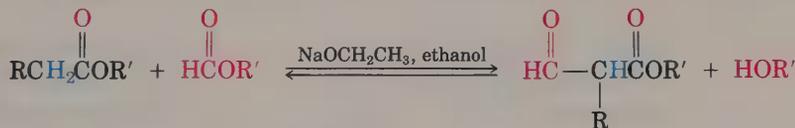
2. Dehydration of aldol products (Section 23.4)



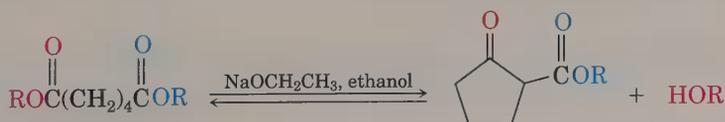
3. Claisen reaction: the condensation between two esters or between one ester and one ketone (Section 23.9)

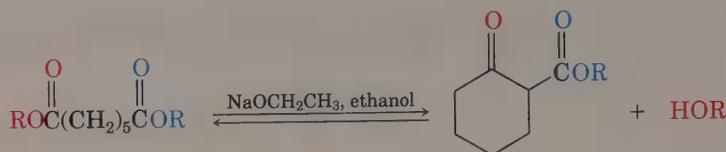


b. Mixed Claisen reaction (Section 23.10)

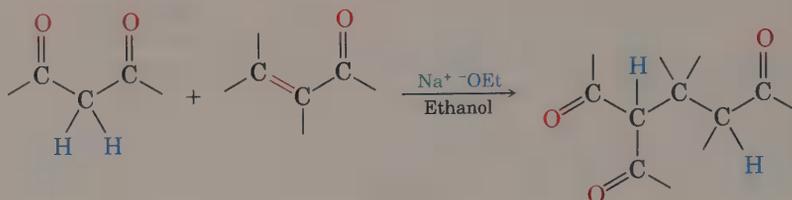


4. Dieckmann cyclization; internal Claisen condensation (Section 23.11)

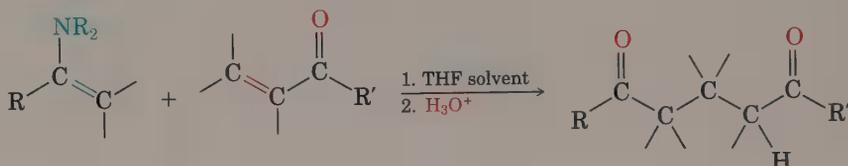




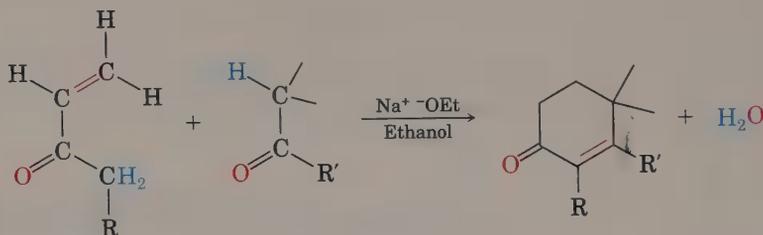
## 5. Michael reaction (Section 23.12)



## 6. Enamine reaction (Section 23.13)

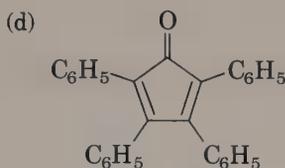
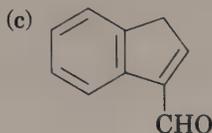


## 7. Robinson annulation reaction (Section 23.14)

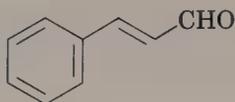


## ADDITIONAL PROBLEMS .....

- 23.27 Which of the following compounds would be expected to undergo aldol self-condensation?
- |                                    |                         |
|------------------------------------|-------------------------|
| (a) Trimethylacetaldehyde          | (b) Cyclobutanone       |
| (c) Benzophenone (diphenyl ketone) | (d) 3-Pentanone         |
| (e) Decanal                        | (f) 3-Phenyl-2-propenal |
- 23.28 Show the product from each compound listed in Problem 23.27 that is capable of undergoing the aldol reaction.
- 23.29 What product would you expect to obtain from aldol cyclization of hexanedial (OHCCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CHO)?
- 23.30 How might you synthesize the following compounds using aldol reactions? In each case, show the structure of the starting ketone(s) or aldehyde(s) you would use.
- |  |                     |
|--|---------------------|
| (a) C <sub>6</sub> H <sub>5</sub> CH=CHCOC <sub>6</sub> H <sub>5</sub> | (b) 2-Cyclohexenone |
|--|---------------------|

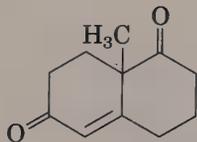


- 23.31** How can you account for the fact that 2,2,6-trimethylcyclohexanone yields no detectable aldol product even though it has an acidic alpha hydrogen?
- 23.32** Cinnamaldehyde, the aromatic constituent of cinnamon oil, can be synthesized by a mixed aldol condensation. Show the starting materials you would use, and formulate the reaction.



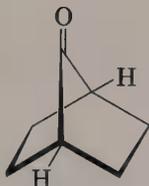
Cinnamaldehyde

- 23.33** The so-called Wieland–Miescher ketone is a valuable starting material used in the laboratory synthesis of steroid hormones. How might you prepare it from 1,3-cyclohexanedione?

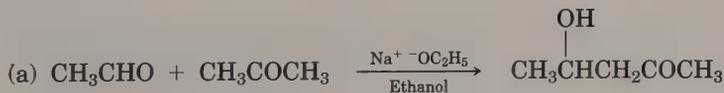


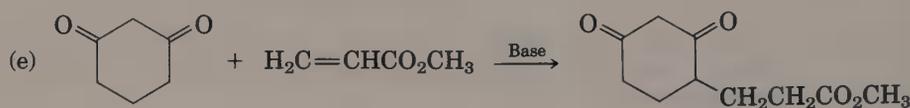
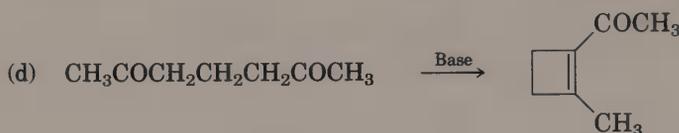
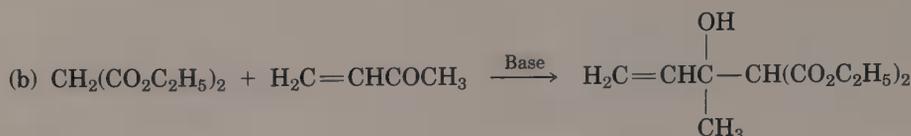
Wieland–Miescher ketone

- 23.34** The bicyclic ketone shown does not undergo aldol self-condensation even though it has two alpha-hydrogen atoms. Explain.



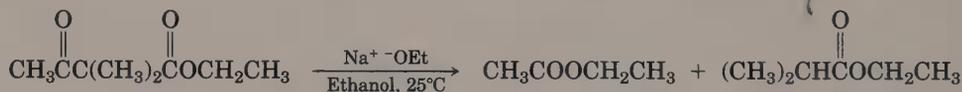
- 23.35** What condensation products would you expect to obtain by treatment of these substances with sodium ethoxide in ethanol?
- (a) Ethyl butanoate (b) Cycloheptanone  
 (c) 3,7-Nonanedione (d) 3-Phenylpropanal
- 23.36** Give the structures of all the possible Claisen condensation products from these reactions. Tell which, if any, you would expect to predominate in each case.
- (a)  $\text{CH}_3\text{CO}_2\text{CH}_3 + \text{CH}_3\text{CH}_2\text{CO}_2\text{CH}_3$  (b)  $\text{C}_6\text{H}_5\text{CO}_2\text{CH}_3 + \text{C}_6\text{H}_5\text{CH}_2\text{CO}_2\text{CH}_3$   
 (c)  $\text{CH}_3\text{OCO}_2\text{CH}_3 + \text{Cyclohexanone}$  (d)  $\text{C}_6\text{H}_5\text{CHO} + \text{CH}_3\text{CO}_2\text{CH}_3$
- 23.37** The following reactions are unlikely to provide the desired product in high yield. What is wrong with each?



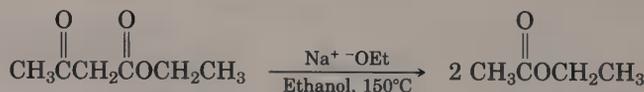


**23.38** In the mixed Claisen reaction of cyclopentanone with ethyl formate, a much higher yield of the desired product is obtained by first mixing the two carbonyl components and then adding base, rather than by first mixing base with cyclopentanone and then adding ethyl formate. Explain.

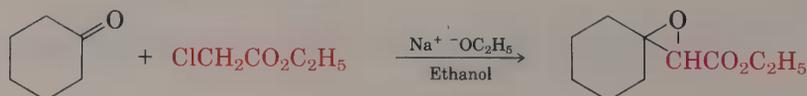
**23.39** Ethyl dimethylacetoacetate reacts instantly at room temperature when treated with ethoxide ion to yield two products, ethyl acetate and ethyl 2-methylpropanoate. Propose a mechanism for this cleavage reaction.



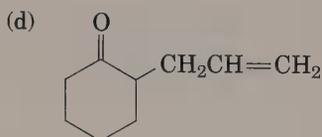
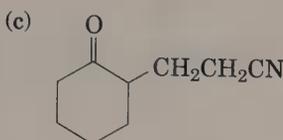
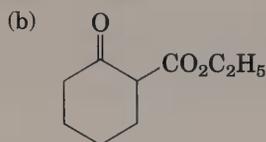
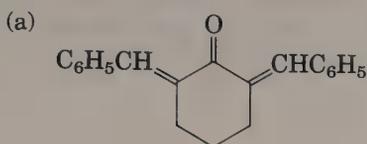
**23.40** In contrast to the rapid reaction shown in Problem 23.39, ethyl acetoacetate itself requires temperatures of over  $150^\circ\text{C}$  to undergo the same kind of cleavage reaction. How can you explain the difference in reactivity?



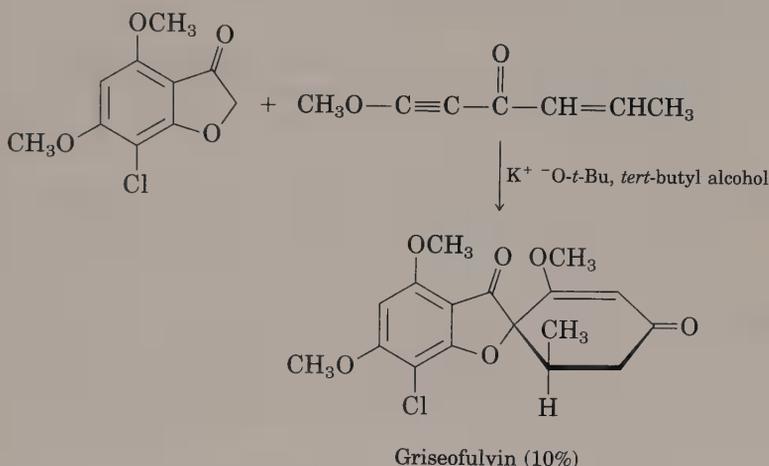
**23.41** The Darzens reaction involves a two-step base-catalyzed condensation of ethyl chloroacetate with a ketone to yield an epoxy ester. The first step is a carbonyl condensation reaction, and the second step is an  $\text{S}_{\text{N}}2$  reaction. Formulate the complete mechanism by which this reaction occurs.



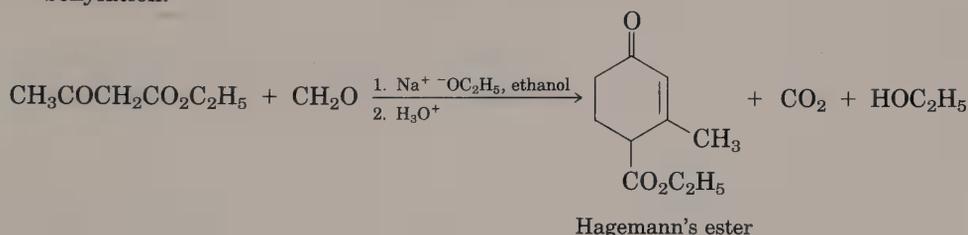
23.42 How would you prepare these compounds from cyclohexanone?



23.43 Griseofulvin, an antibiotic produced by the mold *Penicillium griseofulvum* (Dierckx), has been synthesized by a route that employs a twofold Michael reaction as the key step. Propose a mechanism for this transformation.



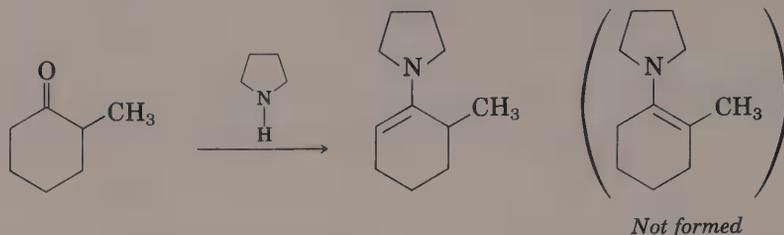
23.44 The compound known as Hagemann's ester is prepared by treatment of a mixture of formaldehyde and ethyl acetoacetate with base, followed by acid-catalyzed decarboxylation:



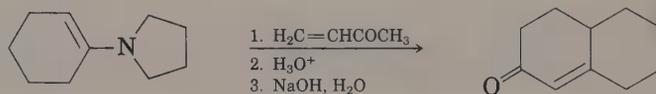
- (a) The first step in the reaction is an aldol-like condensation between ethyl acetoacetate and formaldehyde to yield an  $\alpha,\beta$ -unsaturated product. Write the reaction, and show the structure of the product.
- (b) The second step in the reaction is a Michael reaction between a second equivalent of ethyl acetoacetate and the unsaturated product of step 1. Formulate the reaction, and show the structure of the product.

23.45 The third and fourth steps in the synthesis of Hagemann's ester from ethyl acetoacetate and formaldehyde (Problem 23.44) are an internal aldol cyclization to yield a substituted cyclohexenone ring, and a decarboxylation reaction. Formulate both reactions, and write the products of each step.

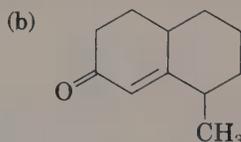
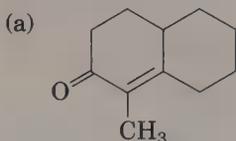
- 23.46 When 2-methylcyclohexanone is converted into an enamine, only one product is formed despite the fact that the starting ketone is unsymmetrical. Build molecular models of the two possible products, and explain the fact that the sole product is the one with the double bond away from the methyl-substituted carbon.



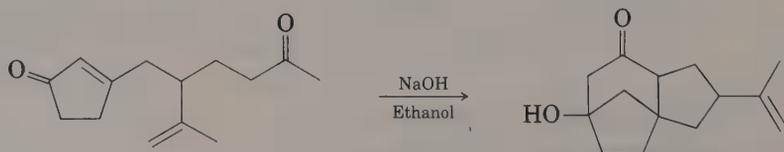
- 23.47 Intramolecular aldol cyclization of 2,5-heptanedione with aqueous NaOH yields a mixture of two enone products in the approximate ratio 9:1. Write their structures, and show how each is formed.
- 23.48 The major product formed by intramolecular aldol cyclization of 2,5-heptanedione (Problem 23.47) has two singlet absorptions in the  $^1\text{H}$  NMR at 1.65  $\delta$  and 1.90  $\delta$ , and has no absorptions in the range 3–10  $\delta$ . What is the structure of this major product?
- 23.49 Treatment of the minor product formed in the intramolecular aldol cyclization of 2,5-heptanedione (Problems 23.47 and 23.48) with aqueous NaOH, converts it into the major product. Propose a mechanism to account for this base-catalyzed isomerization.
- 23.50 The Stork enamine reaction and the intramolecular aldol reaction can be carried out in sequence to allow the synthesis of cyclohexenone rings. For example, reaction of the pyrrolidine enamine of cyclohexanone with 3-buten-2-one, followed by enamine hydrolysis and base treatment, yields the product indicated. Show the mechanisms of the different steps.



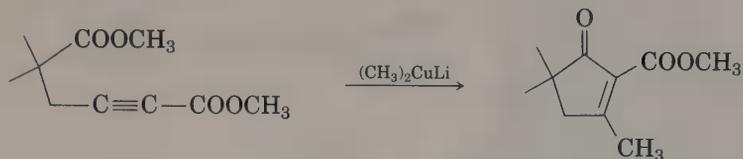
- 23.51 How could you prepare these cyclohexenones by combining a Stork enamine reaction with an intramolecular aldol condensation? [Hint: See Problem 23.46.]



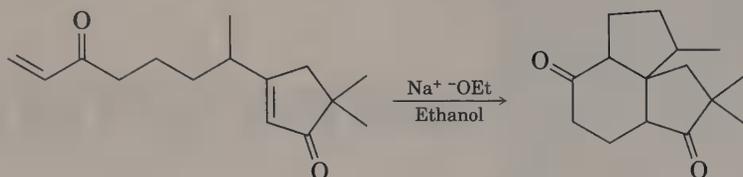
- 23.52 Propose a mechanism to account for the following reaction:



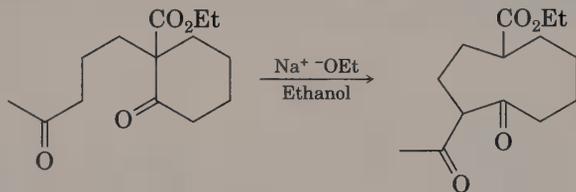
23.53 Propose a mechanism to account for the following reaction:

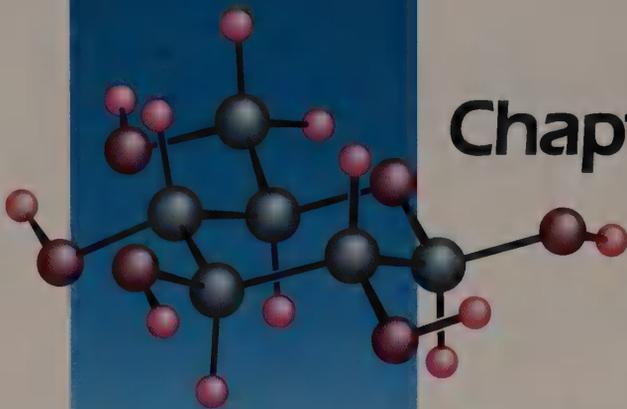


23.54 Propose a mechanism to account for the following reaction:



23.55 Propose a mechanism to account for the following reaction:



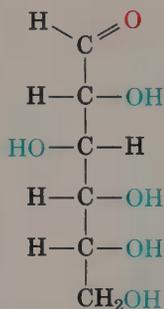


# Chapter 24

## Carbohydrates

Carbohydrates are everywhere in nature. They occur in every living organism and are essential to life. The sugar and starch in food, and the cellulose in wood, paper, and cotton, are nearly pure carbohydrate. Modified carbohydrates form part of the coating around living cells; other carbohydrates are found in the DNA that carries genetic information; and still others, such as gentamicin, are invaluable as medicines.

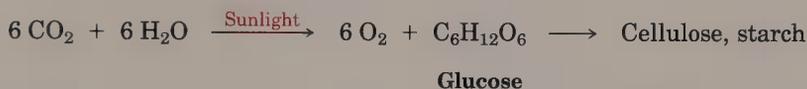
The word **carbohydrate** derives historically from the fact that glucose, the first simple carbohydrate to be purified, has the molecular formula  $C_6H_{12}O_6$  and was originally thought to be a “hydrate of carbon,”  $C_6(H_2O)_6$ . This view was soon abandoned, but the name persisted. Today, the term *carbohydrate* is used to refer loosely to the broad class of polyhydroxylated aldehydes and ketones commonly called *sugars*.



Glucose (also called dextrose),  
 a pentahydroxypentanal

Carbohydrates are synthesized by green plants during photosynthesis, a complex process during which carbon dioxide is converted into glucose.

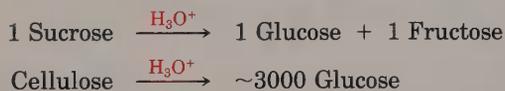
Many molecules of glucose are then chemically linked for storage by the plant in the form of either cellulose or starch. It has been estimated that more than 50% of the dry weight of the earth's biomass—all plants and animals—consists of glucose polymers. When eaten and then metabolized, carbohydrates provide the major source of energy required by organisms. Thus, carbohydrates act as the chemical intermediaries by which solar energy is stored and used to support life.



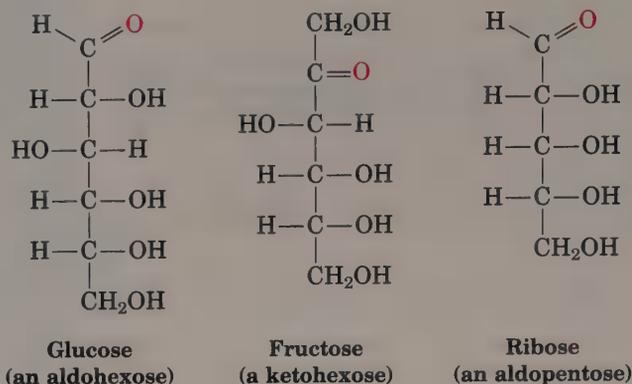
When eaten, glucose can be either metabolized in the body to provide immediate energy or stored by the body in the form of glycogen for use at a later time. Since humans and most other mammals lack the enzymes needed for digestion of cellulose, they require starch as their dietary source of carbohydrates. Grazing animals such as cows, however, contain in their rumen microorganisms that are able to digest cellulose. The energy stored in cellulose is thus moved up the biological food chain when these animals are used for food.

## 24.1 Classification of Carbohydrates

Carbohydrates are generally classed into two groups: simple and complex. **Simple sugars**, or **monosaccharides**, are carbohydrates like glucose and fructose that can't be hydrolyzed into smaller molecules. **Complex carbohydrates** are made of two or more simple sugars linked together. For example, sucrose (table sugar) is a **disaccharide** (two sugars) made up of one glucose molecule linked to one fructose molecule. Similarly, cellulose is a polysaccharide (many sugars) made up of several thousand glucose molecules linked together. Hydrolysis of these polysaccharides breaks them down into their constituent monosaccharide units.

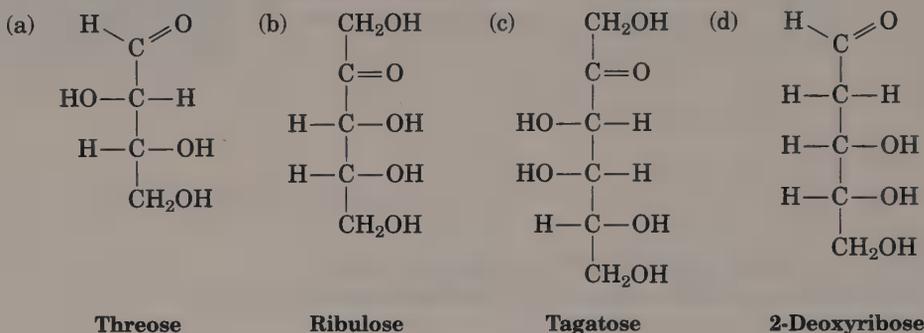


Monosaccharides can be further classified as either **aldoses** or **ketoses**. The *-ose* suffix is used to designate a carbohydrate, and the *aldo-* and *keto-* prefixes designate the nature of the carbonyl group (aldehyde or ketone). The number of carbon atoms in the monosaccharide is given by using *tri-*, *tetr-*, *pent-*, *hex-*, and so forth as the parent name. For example, glucose is an *aldohexose*, a six-carbon aldehyde sugar; fructose is a *ketohexose*, a six-carbon keto sugar; and ribose is an *aldopentose*, a five-carbon aldehyde sugar. Most of the commonly occurring sugars are either aldopentoses or aldohexoses.



**PROBLEM** .....

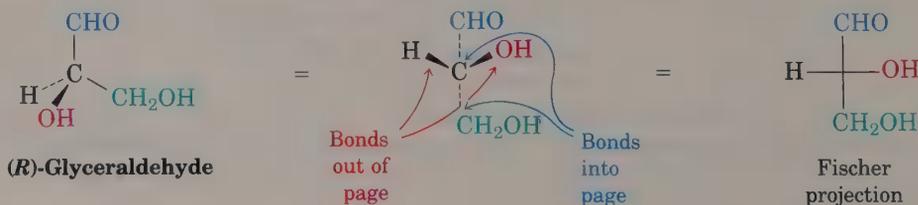
24.1 Classify each of the following monosaccharides:



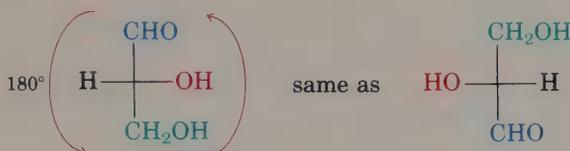
## 24.2 Fischer Projections for Depicting Carbohydrates

Since all carbohydrates have chiral carbon atoms, it was recognized long ago that a standard method of representation was needed to designate carbohydrate stereochemistry. The method most commonly used employs **Fischer projections** (Section 9.13) for depicting a stereogenic center on a flat page.

Recall that a tetrahedral carbon atom is represented in a Fischer projection by two crossed lines. By convention, the horizontal lines represent bonds coming out of the page, and the vertical lines represent bonds going into the page. For example, (*R*)-glyceraldehyde, the simplest monosaccharide, can be drawn as follows:

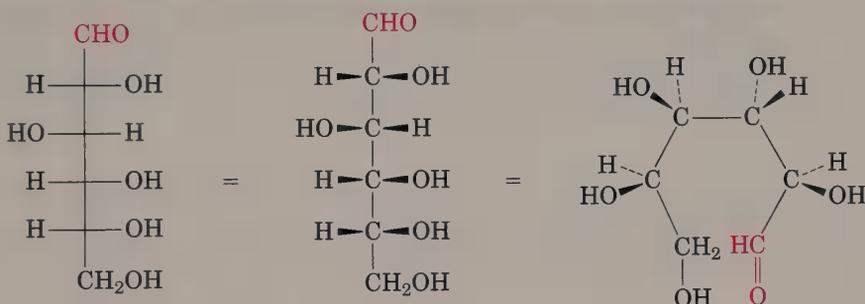


Recall also that Fischer projections can be rotated on the page by  $180^\circ$  without changing their meaning, but not by  $90^\circ$  or  $270^\circ$ :



(*R*)-Glyceraldehyde

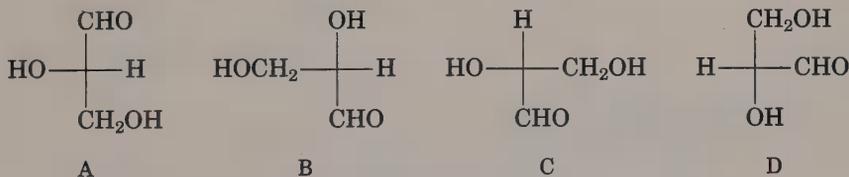
Fischer projections can be used to depict more than one stereogenic center in a molecule simply by stacking the stereogenic centers one on top of the other. By convention, the carbonyl carbon is always placed either at the top or near the top when showing the Fischer projection of a carbohydrate. For example, glucose has four stereogenic centers stacked on top of each other in Fischer projection:



Glucose  
(carbonyl group at top)

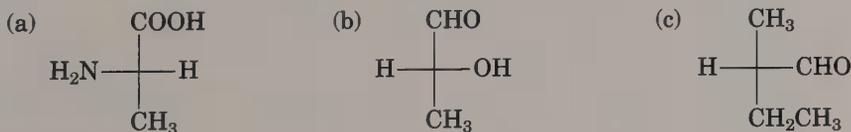
PROBLEM .....

24.2 Which of these Fischer projections of glyceraldehyde represent the same enantiomer?



PROBLEM .....

24.3 Convert these Fischer projections into tetrahedral representations, and assign *R* or *S* stereochemistry to each.

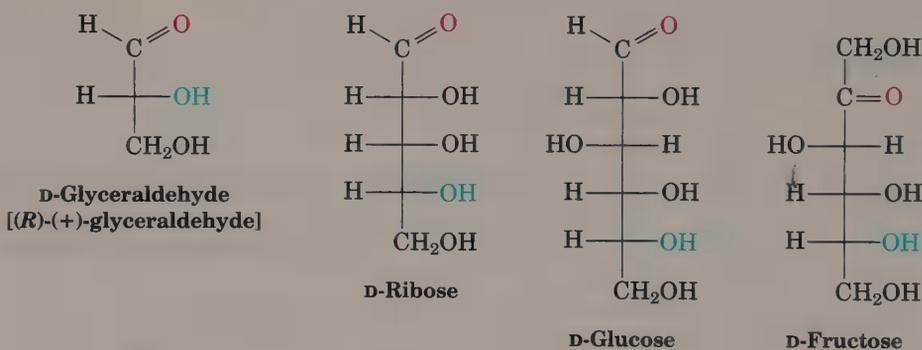


## 24.3 D,L Sugars

Glyceraldehyde has one chiral carbon atom and can therefore have two enantiomeric, mirror-image forms. Only one of these enantiomers occurs naturally, however, and this natural enantiomer is dextrorotatory. That is, a sample of naturally occurring glyceraldehyde placed in a polarimeter will rotate plane-polarized light in a clockwise direction, denoted (+).

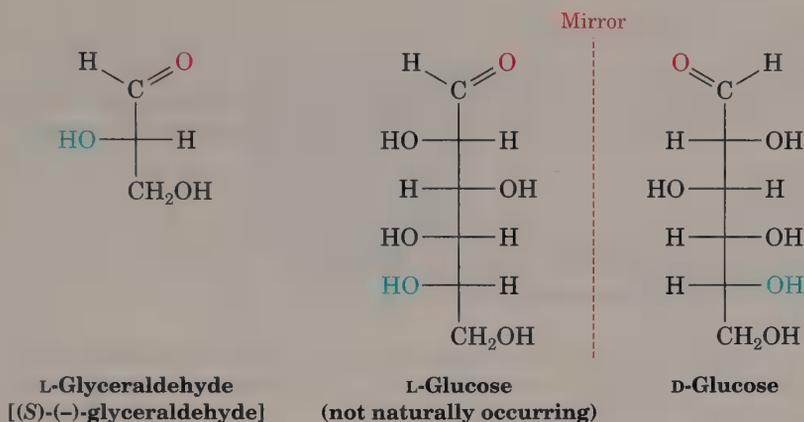
Since (+)-glyceraldehyde is known to have the *R* configuration at C2, it can be represented in Fischer projection as shown in Figure 24.1. For historical reasons dating back long before the adoption of the *R,S* system, (*R*)-(+)-glyceraldehyde is also referred to as *D-glyceraldehyde* (*D* from dextrorotatory). The nonnatural enantiomer (*S*)-(–)-glyceraldehyde is similarly known as *L-glyceraldehyde* (*L* from levorotatory).

Because of the way monosaccharides are biosynthesized in nature, glucose, fructose, and almost all other naturally occurring monosaccharides have the same stereochemical configuration as *D*-glyceraldehyde at the chiral carbon atom farthest from the carbonyl group. In Fischer projections, therefore, most naturally occurring sugars have the hydroxyl group at the lowest chiral carbon atom pointing to the right (Figure 24.1). Such compounds are referred to as **D sugars**.



**Figure 24.1** Some naturally occurring *D* sugars. The hydroxyl group at the stereogenic center farthest from the carbonyl group is on the right when the molecule is drawn in Fischer projection.

In contrast to **D sugars**, all **L sugars** have the hydroxyl group at the lowest chiral carbon atom on the *left* in Fischer projection. Thus, **L sugars** are mirror images (enantiomers) of **D sugars**. Note that the **D** and **L** notations have no relation to the direction in which a given sugar rotates plane-polarized light; a **D** sugar may be either dextrorotatory or levorotatory. The prefix **D** indicates only that the stereochemistry of the lowest chiral carbon atom is to the right in Fischer projection when the molecule is drawn in the standard way with the carbonyl group at or near the top.

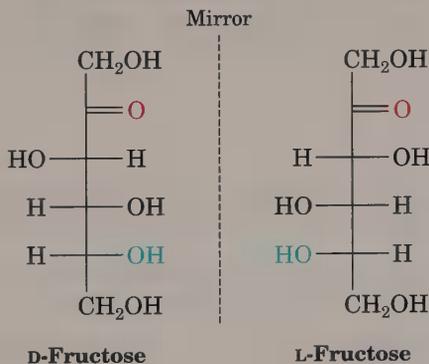


The D,L system of carbohydrate nomenclature is of limited use, since it describes the configuration at only one stereogenic center and says nothing about other stereogenic centers that may be present. The advantage of the system, though, is that it allows a person to relate one sugar to another rapidly and visually.

**PRACTICE PROBLEM**.....

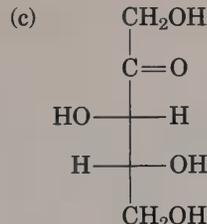
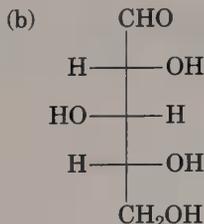
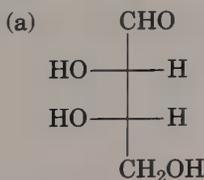
Draw a Fischer projection of L-fructose.

**Solution** Since L-fructose is the enantiomer (mirror image) of D-fructose, we simply look at the structure of D-fructose and then reverse the configuration at each stereogenic center.



**PROBLEM**.....

**24.4** Assign *R* or *S* configuration to each chiral carbon atom in the following sugars, and tell if each is a D sugar or an L sugar.



## PROBLEM.....

- 24.5 (+)-Arabinose, an aldopentose that is widely distributed in plants, can be named systematically as (2*R*,3*S*,4*S*)-2,3,4,5-tetrahydroxypentanal. Draw a Fischer projection of (+)-arabinose, and identify it as a *D* sugar or an *L* sugar.
- .....

## 24.4 Configurations of the Aldoses

---

Aldotetroses are four-carbon sugars with two stereogenic centers. There are  $2^2 = 4$  possible stereoisomeric aldotetroses, or two *D,L* pairs of enantiomers, called *erythrose* and *threose*.

Aldopentoses have three stereogenic centers, leading to a total of  $2^3 = 8$  possible stereoisomers, or four *D,L* pairs of enantiomers. These four pairs are called *ribose*, *arabinose*, *xylose*, and *lyxose*. All except lyxose occur widely in nature. *D*-Ribose is an important constituent of RNA (ribonucleic acid); *L*-arabinose is found in many plants; and *D*-xylose is found in wood.

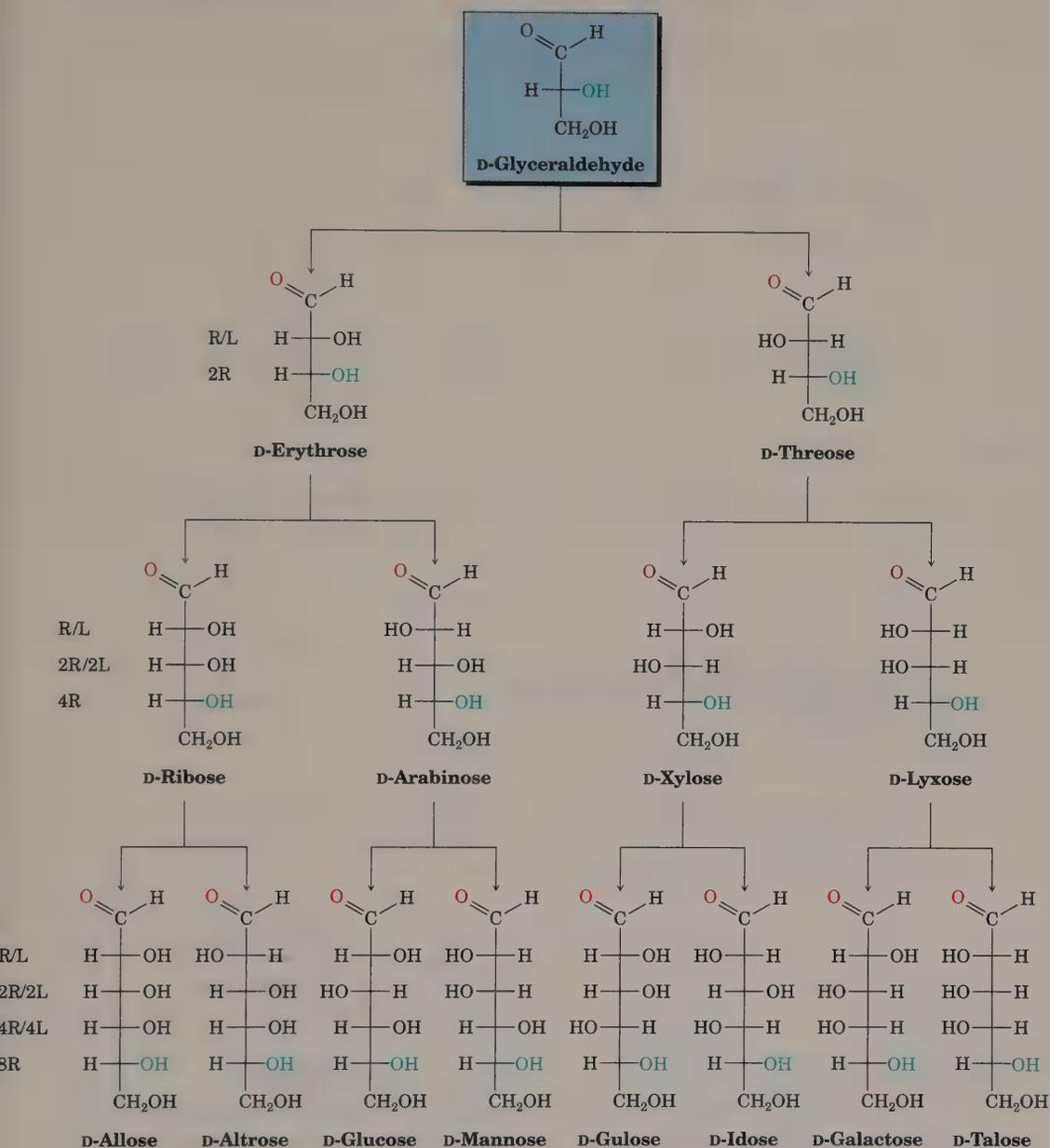
Aldohexoses have four stereogenic centers, for a total of  $2^4 = 16$  possible stereoisomers, or eight *D,L* pairs of enantiomers. The names of the eight are *allose*, *altrose*, *glucose*, *mannose*, *gulose*, *idose*, *galactose*, and *talose*. Of the eight, only *D*-glucose (from starch and cellulose) and *D*-galactose (from gums and fruit pectins) are found widely in nature. *D*-Mannose and *D*-talose also occur naturally, but in lesser abundance.

Fischer projections of the four-, five-, and six-carbon aldoses are shown in Figure 24.2 for the *D* series. Starting from *D*-glyceraldehyde, we can construct the two *D*-aldotetroses by inserting a new chiral carbon atom just below the aldehyde carbon. Each of the two *D*-aldotetroses then leads to two *D*-aldopentoses (four total), and each of the four *D*-aldopentoses leads to two *D*-aldohexoses (eight total).

Louis Fieser<sup>1</sup> of Harvard suggested this procedure for remembering the names and structures of the eight *D*-aldohexoses:

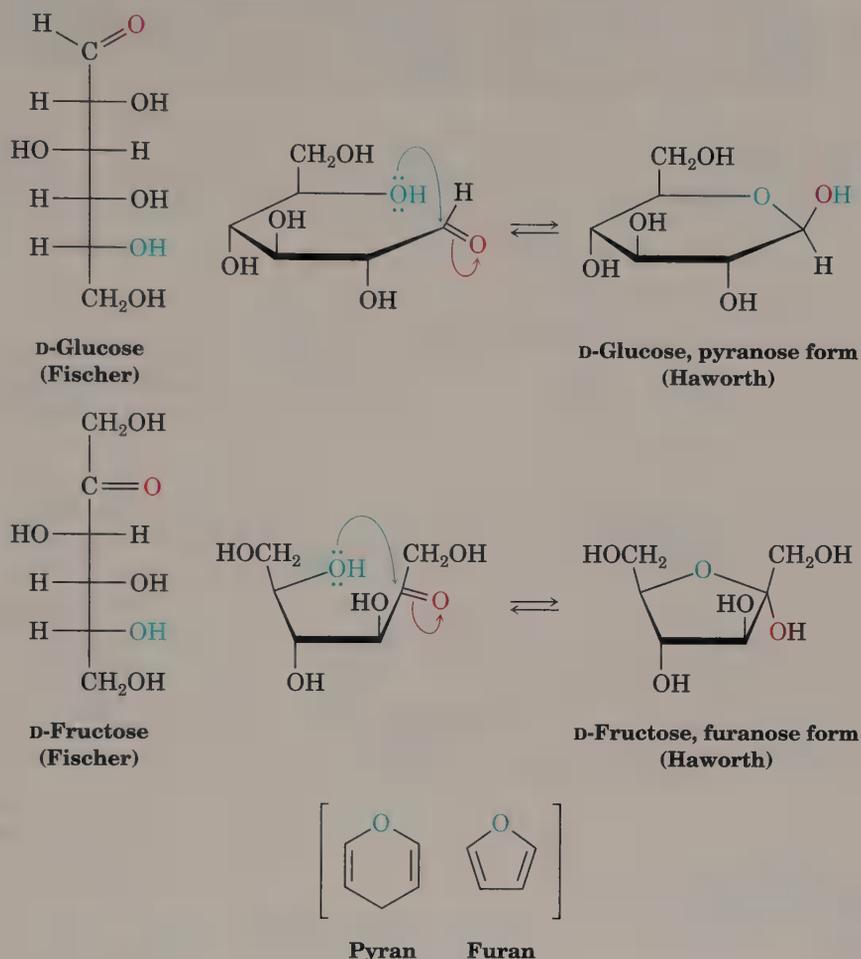
1. Set up eight Fischer projections with the aldehyde group on top and the  $\text{CH}_2\text{OH}$  group at the bottom.
2. Indicate stereochemistry at C5 by placing all eight hydroxyl groups to the right (*D* series).
3. Indicate stereochemistry at C4 by alternating four hydroxyl groups to the right and four to the left.
4. Indicate stereochemistry at C3 by alternating two hydroxyl groups to the right, two to the left, and so on.
5. Indicate stereochemistry at C2 by alternating hydroxyl groups right, left, right, left, and so on.

<sup>1</sup>Louis F. Fieser (1899–1977); b. Columbus, Ohio; Ph.D. Harvard (Conant); professor, Bryn Mawr College, Harvard University.



**Figure 24.2** Configurations of D aldoses. The structures are arranged in order from left to right so that the hydroxyl groups on C2 alternate right/left (R/L) in going across a series. Similarly, the hydroxyl groups at C3 alternate two right/two left (2R/2L); the hydroxyl groups at C4 alternate 4R/4L; and the hydroxyl groups at C5 are to the right in all eight (8R).





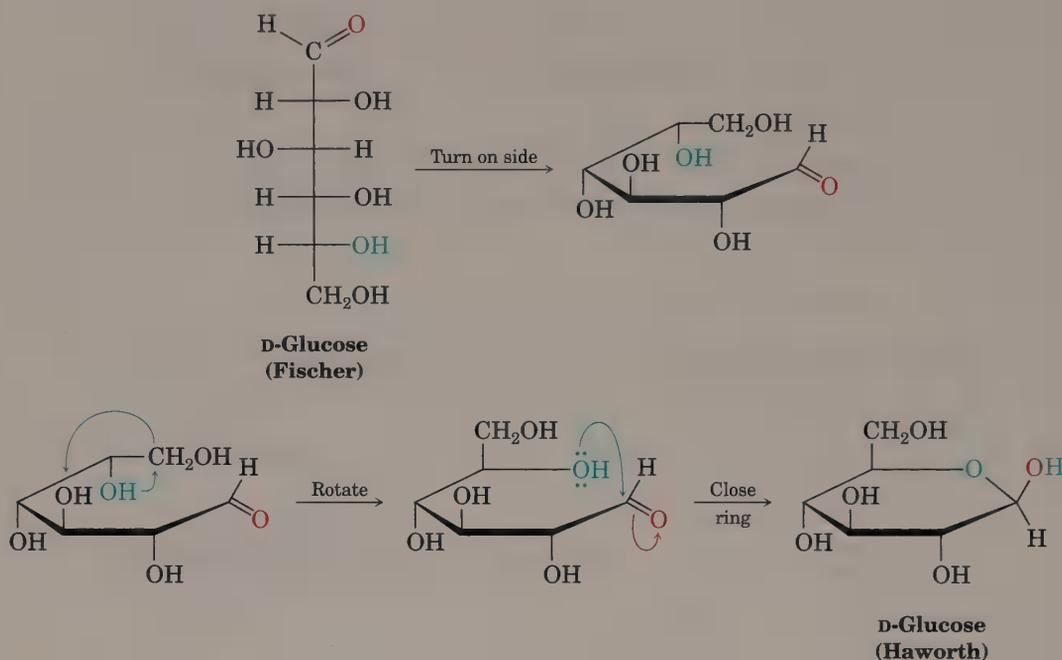
**Figure 24.3** Glucose and fructose in their cyclic furanose and pyranose forms.

Pyranose and furanose rings are often represented using the **Haworth<sup>2</sup> projections** shown in Figure 24.3, rather than Fischer projections. In a Haworth projection, the hemiacetal ring is drawn as if it were flat and is viewed edge-on with the oxygen atom at the upper right. Though convenient, this view is not really accurate because pyranose rings are actually chair shaped like cyclohexane (Section 4.9), rather than flat. Nevertheless, Haworth projections are widely used because they make it possible to see at a glance the *cis*–*trans* relationships among hydroxyl groups on the ring.

When converting from one kind of projection to the other, remember that a hydroxyl on the *right* in a Fischer projection is *down* in a Haworth

<sup>2</sup>Sir Walter Norman Haworth (1883–1950); b. Chorley, Lancashire; Ph.D. Göttingen; D.Sc. Manchester; professor, University of Birmingham; Nobel Prize (1937).

projection. Conversely, a hydroxyl on the *left* in a Fischer projection is *up* in a Haworth projection. For *D* sugars, the terminal  $-\text{CH}_2\text{OH}$  group is always up in Haworth projections, whereas for *L* sugars the  $-\text{CH}_2\text{OH}$  group is down. Figure 24.4 illustrates the conversion for *D*-glucose.

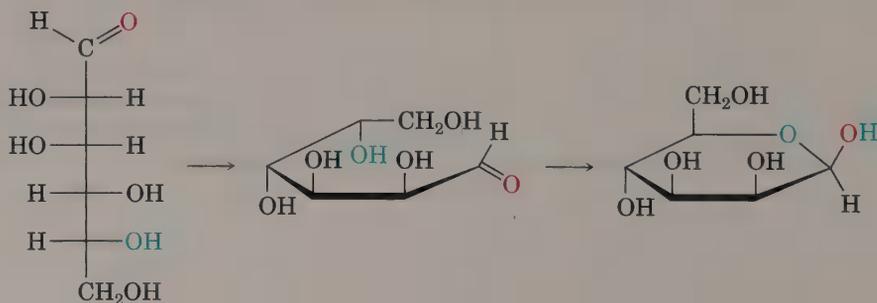


**Figure 24.4** Interconversion of Fischer and Haworth projections of *D*-glucose.

**PRACTICE PROBLEM.** .....

*D*-Mannose differs from *D*-glucose in its stereochemistry at C2. Draw a Haworth projection of *D*-mannose in its pyranose form.

**Solution** First draw a Fischer projection of *D*-mannose. Then lay it on its side, and curl it around so that the aldehyde group (C1) is toward the front and the  $\text{CH}_2\text{OH}$  group (C6) is toward the rear. Now connect the hydroxyl at C5 to the C1 carbonyl group to form a pyranose ring.



PROBLEM.....

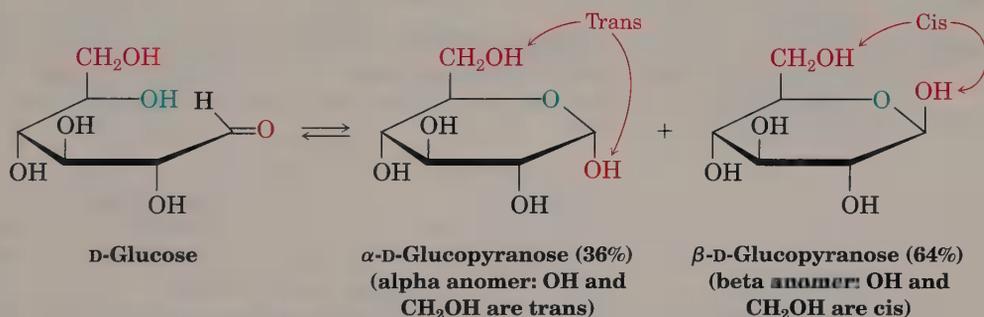
- 24.9 D-Galactose differs from D-glucose in its stereochemistry at C4. Draw a Haworth projection of D-galactose in its pyranose form.

PROBLEM.....

- 24.10 Draw Haworth projections of L-glucose in its pyranose form and D-ribose in its furanose form.

## 24.6 Monosaccharide Anomers: Mutarotation

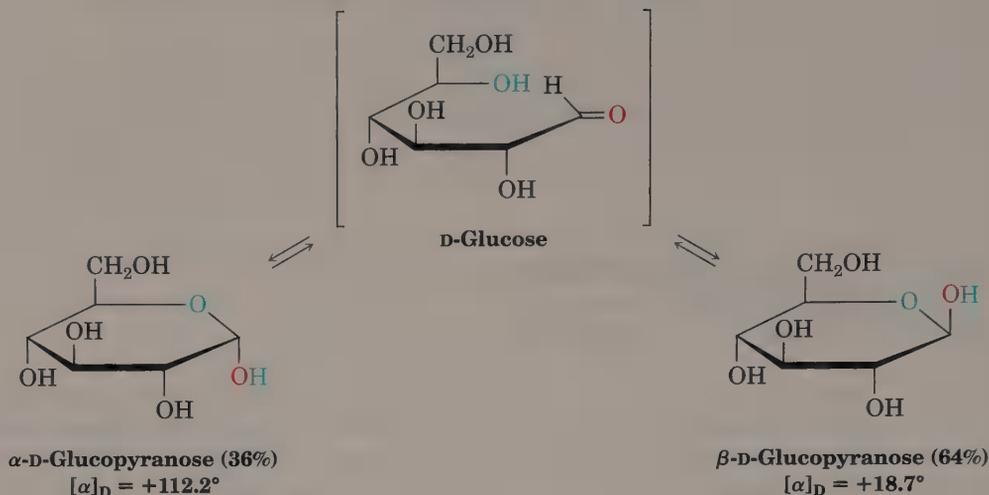
When an open-chain monosaccharide cyclizes to a furanose or pyranose form, a new stereogenic center is formed at what used to be the carbonyl carbon. The two diastereomers produced are called **anomers**, and the hemiacetal carbon atom is referred to as the **anomeric center**. For example, glucose cyclizes reversibly in aqueous solution to a 36:64 mixture of two anomers. The minor anomer with the C1 -OH group trans to the -CH<sub>2</sub>OH substituent at C5 (down in a Haworth projection) is called the **alpha anomer**;<sup>3</sup> its complete name is  $\alpha$ -D-glucopyranose. The major anomer with the C1 -OH group cis to the -CH<sub>2</sub>OH substituent at C5 (up in a Haworth projection) is called the **beta anomer**; its complete name is  $\beta$ -D-glucopyranose.



Both anomers of D-glucopyranose can be crystallized and purified. Pure  $\alpha$ -D-glucopyranose has a melting point of 146°C and a specific rotation,  $[\alpha]_D$ , of +112.2°; pure  $\beta$ -D-glucopyranose has a melting point of 148–155°C and a specific rotation of +18.7°. When a sample of either pure anomer is dissolved in water, however, the optical rotations slowly change and ultimately converge to a constant value of +52.6°. The specific rotation of the alpha-anomer solution decreases from +112.2° to +52.6°, and the specific rotation of the beta-anomer solution increases from +18.7° to +52.6°. Known as **mutarotation**, this phenomenon is due to the slow conversion of the pure  $\alpha$  and  $\beta$  anomers into the 36:64 equilibrium mixture.

<sup>3</sup>An alpha anomer is formally defined as having the same configuration at the anomeric center and at the last asymmetric center. A beta anomer has opposite configurations at the anomeric center and the last asymmetric center.

Mutarotation occurs by a reversible ring-opening of each anomer to the open-chain aldehyde, followed by reclosure. Although equilibration is slow at neutral pH, it is catalyzed by both acid and base.



PROBLEM .....

24.11 Draw the two anomers of D-fructose in their furanose forms.

PROBLEM .....

24.12 Knowing that the specific rotation of pure  $\alpha$ -D-glucopyranose is  $+112.2^\circ$  and that the specific rotation of pure  $\beta$ -D-glucopyranose is  $+18.7^\circ$ , show how the equilibrium percentages of alpha and beta anomers can be calculated from the equilibrium specific rotation of  $+52.6^\circ$ .

PROBLEM .....

24.13 Many other sugars besides glucose exhibit mutarotation. For example,  $\alpha$ -D-galactopyranose has  $[\alpha]_D = +150.7^\circ$ , and  $\beta$ -D-galactopyranose has  $[\alpha]_D = +52.8^\circ$ . If either anomer is dissolved in water and allowed to reach equilibrium, the specific rotation of the solution is  $+80.2^\circ$ . What are the percentages of each anomer at equilibrium? Draw the pyranose forms of both anomers using Haworth projections.

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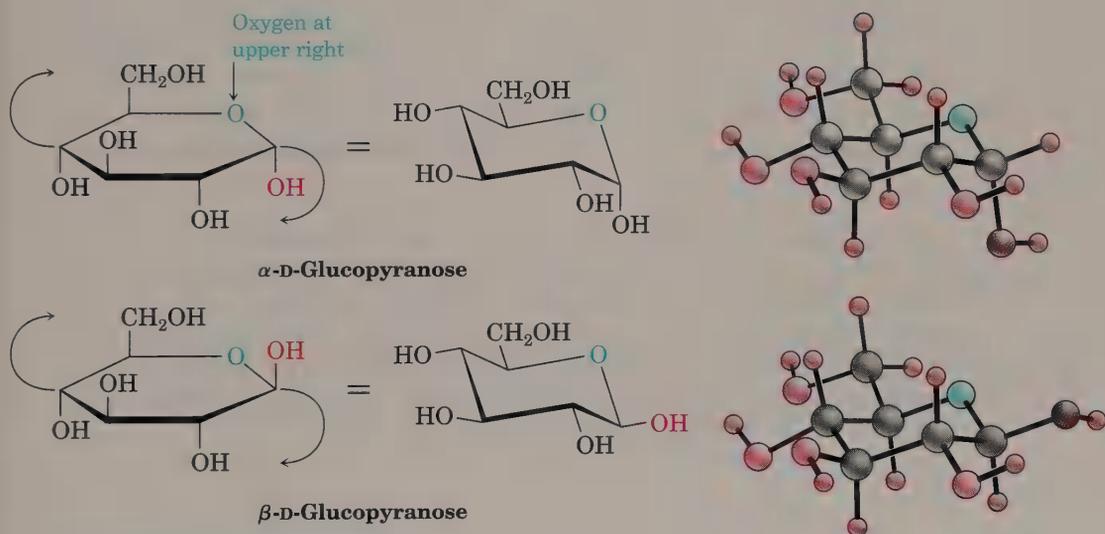
## 24.7 Conformations of Monosaccharides

Although Haworth projections are fairly easy to draw and readily show cis-trans relationships between substituents on furanose and pyranose rings, they don't give an accurate three-dimensional picture of molecular conformation. Pyranose rings, like cyclohexane rings (Section 4.9), have a chair-like geometry with axial and equatorial substituents. Any substituent that

is up in a Haworth projection is also up in a chair conformation, and any substituent that is down in a Haworth projection is down in the chair conformation. Haworth projections can be converted into chair representations by following three steps:

1. Draw the Haworth projection with the ring oxygen atom at the upper right.
2. Raise the leftmost carbon atom (C4) *above* the ring plane.
3. Lower the anomeric carbon atom (C1) *below* the ring plane.

Figure 24.5 shows how this conversion is done for  $\alpha$ -D-glucopyranose and  $\beta$ -D-glucopyranose. Make molecular models to see the process more clearly.



**Figure 24.5** Chair representations of  $\alpha$ -D-glucopyranose and  $\beta$ -D-glucopyranose.

Note that in  $\beta$ -D-glucopyranose all the substituents on the ring are equatorial. Thus,  $\beta$ -D-glucopyranose is the least sterically crowded and most stable of the eight D-aldohexoses.

**PROBLEM** .....

- 24.14** Draw chair conformations of  $\beta$ -D-galactopyranose and  $\beta$ -D-mannopyranose. Label the ring substituents as either axial or equatorial. Which would you expect to be more stable: galactose or mannose?

**PROBLEM** .....

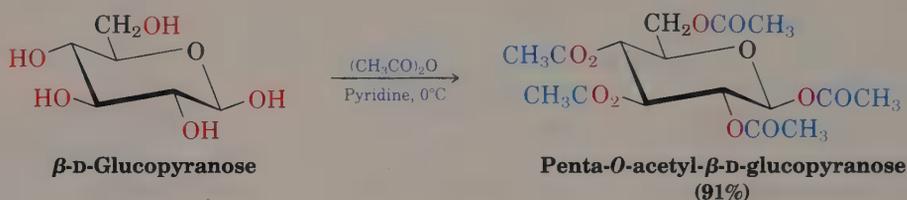
- 24.15** Draw a chair conformation of  $\beta$ -L-glucopyranose, and label the substituents as either axial or equatorial.
- .....

## 24.8 Reactions of Monosaccharides

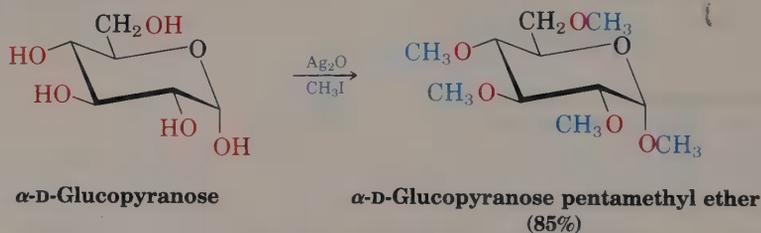
### ESTER AND ETHER FORMATION

Monosaccharides behave as simple alcohols in much of their chemistry. For example, carbohydrate hydroxyl groups can be converted into esters and ethers.

Esterification is normally carried out by treating the carbohydrate with an acid chloride or acid anhydride in the presence of a base. All the hydroxyl groups react, including the anomeric one. For example,  $\beta$ -D-glucopyranose is converted into its pentaacetate by treatment with acetic anhydride in pyridine solution:



Carbohydrates can be converted into ethers by treatment with an alkyl halide in the presence of base (the Williamson ether synthesis, Section 18.4). Normal Williamson conditions using a strong base tend to degrade the sensitive sugar molecules, but Purdie<sup>4</sup> showed in 1903 that silver oxide works particularly well and that high yields of ethers are obtained. For example,  $\alpha$ -D-glucopyranose is converted into its pentamethyl ether in 85% yield on reaction with iodomethane and silver oxide:

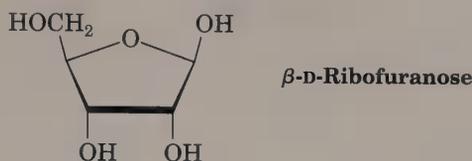


Ester and ether derivatives of carbohydrates are often prepared because they are easier to work with than the free sugars are. Because of their many hydroxyl groups, monosaccharides are usually soluble in water but insoluble in organic solvents such as ether. They are also difficult to purify and have a tendency to form syrups rather than crystals when water is removed. Ester and ether derivatives, however, behave like most other organic compounds in that they tend to be soluble in organic solvents and to be readily purified and crystallized.

<sup>4</sup>Thomas Purdie (1843–1916); b. Biggar, Scotland; Ph.D. Würzburg; professor, St. Andrews University.

## PROBLEM.....

- 24.16 Draw the products you would obtain by reaction of  $\beta$ -D-ribofuranose with:  
 (a)  $\text{CH}_3\text{I}$ ,  $\text{Ag}_2\text{O}$     (b)  $(\text{CH}_3\text{CO})_2\text{O}$ , pyridine

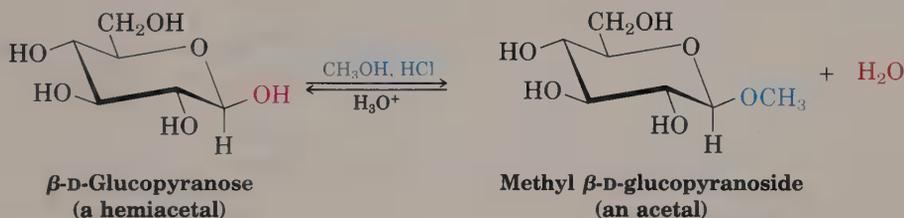


## GLYCOSIDE FORMATION

We saw in Section 19.14 that treatment of a hemiacetal with an alcohol and an acid catalyst yields an acetal:

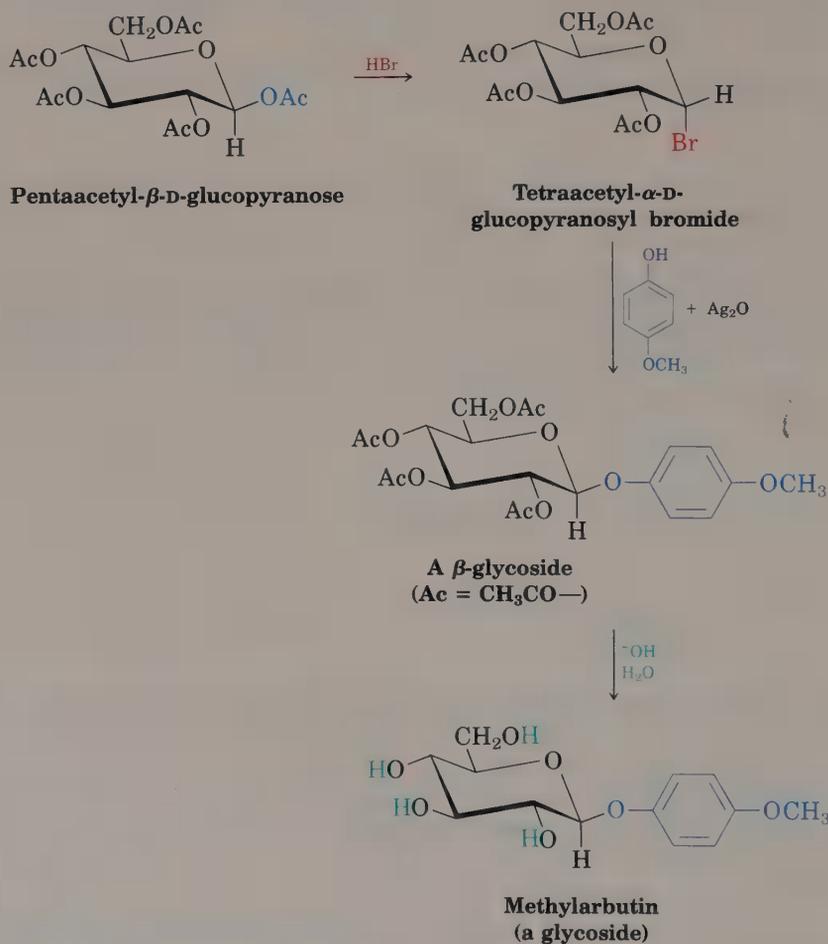
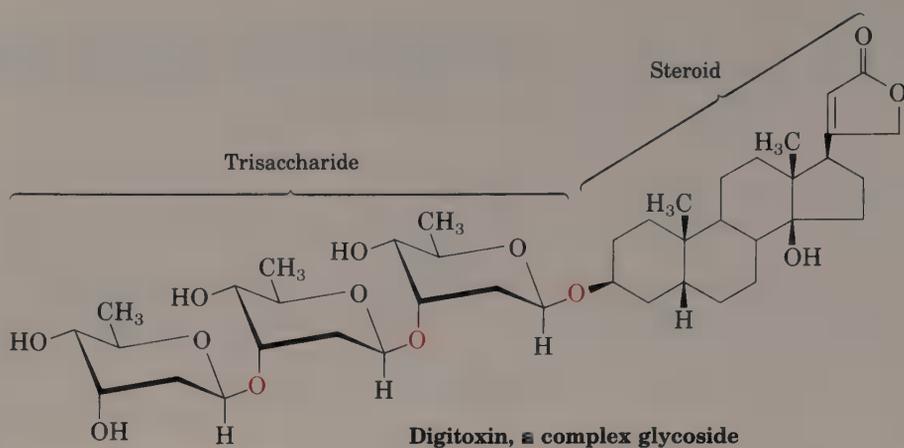


In the same way, treatment of a monosaccharide hemiacetal with an alcohol and an acid catalyst yields an acetal in which the anomeric hydroxyl has been replaced by an alkoxy group. For example, reaction of glucose with methanol gives methyl  $\beta$ -D-glucopyranoside:



Carbohydrate acetals are called **glycosides**. They are named by citing the alkyl group and adding the *-oside* suffix to the name of the specific sugar. Note that glycosides, like all acetals, are stable to water. They aren't in equilibrium with an open-chain form, and they don't show mutarotation. They can, however, be converted back to the free monosaccharide by hydrolysis with aqueous acid.

Glycosides are widespread in nature, and a great many biologically important molecules contain glycosidic linkages. For example, digitoxin, the active component of the digitalis preparations used for treatment of heart disease, is a glycoside consisting of a complex steroid alcohol linked to a trisaccharide. Note also that the three sugars are linked to each other by glycosidic bonds.



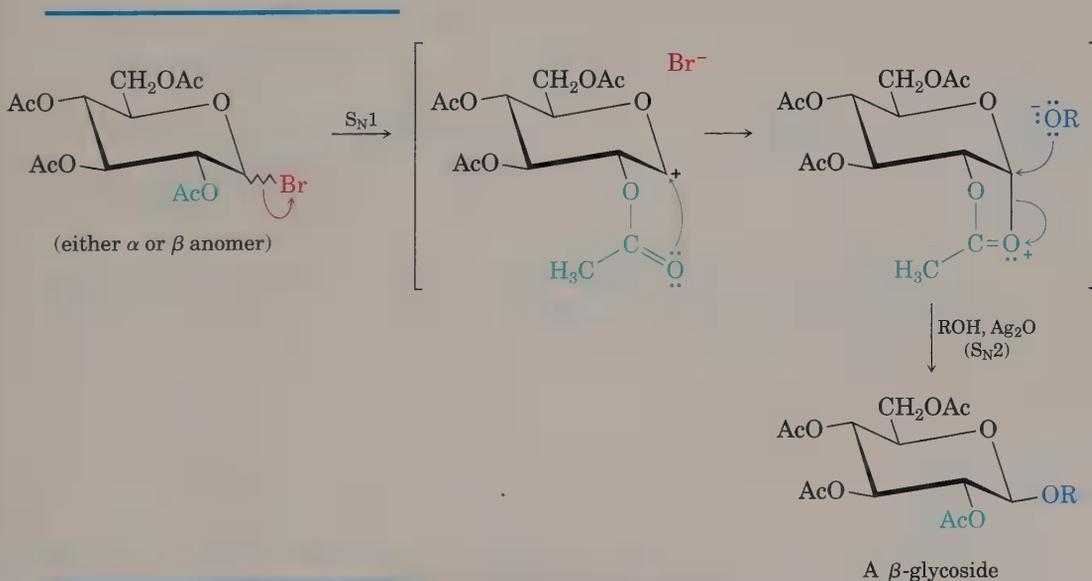
**Figure 24.6** Synthesis of the glycoside methylarbutin by Koenigs–Knorr reaction of pentaacetyl- $\beta$ -D-glucopyranose with *p*-methoxyphenol.

The laboratory synthesis of glycosides is often difficult, and a successful reaction scheme depends on the structures of both alcohol and saccharide components. One method (the **Koenigs–Knorr**<sup>5</sup> reaction) that is particularly suitable for preparing glucose  $\beta$ -glycosides involves treatment of glucose pentaacetate with HBr, followed by addition of the appropriate alcohol in the presence of silver oxide.

The Koenigs–Knorr reaction sequence involves formation of a pyranosyl bromide, followed by nucleophilic substitution, and yields a  $\beta$ -glycoside. For example, methylarbutin, a glycoside found in pear leaves, has been prepared by reaction between *p*-methoxyphenol and tetraacetyl- $\alpha$ -D-glucopyranosyl bromide (Figure 24.6).

Although the Koenigs–Knorr reaction appears to involve a simple back-side  $S_N2$  displacement of bromide ion by alkoxide ion in a Williamson ether synthesis, the actual situation is more complex. Both alpha and beta anomers of tetraacetyl-D-glucopyranosyl bromide give the *same*  $\beta$ -glycoside product, suggesting that both anomers react by a common pathway.

The results are best explained by assuming that tetraacetyl-D-glucopyranosyl bromide (either alpha or beta anomer) undergoes a spontaneous  $S_N1$  loss of bromide ion, followed by internal reaction of the cationic center at C1 with an oxygen atom of the ester group at C2 to form a stabilized oxonium ion intermediate. Since the ester group at C2 is on the bottom of the glucose ring, the new carbon–oxygen bond also forms from the bottom. An  $S_N2$  displacement of the oxonium ion by back-side attack at C1 then occurs with the usual inversion of configuration, yielding a  $\beta$ -glycoside and regenerating the acetate ester group at C2 (Figure 24.7).



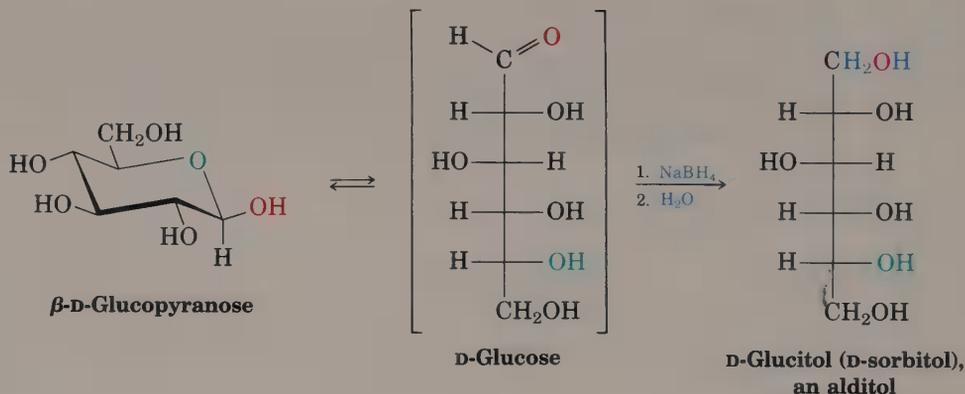
**Figure 24.7** Mechanism of the Koenigs–Knorr reaction, showing the neighboring-group effect of a nearby acetoxy.

<sup>5</sup>Ludwig Knorr (1859–1921); Ph.D. Erlangen, 1882; professor, University of Jena.

The kind of participation shown by the nearby acetoxyl group in the Koenigs–Knorr reaction is referred to as a **neighboring-group effect** and is a common occurrence in organic chemistry. Neighboring-group effects are usually noticeable only because they affect the rate or stereochemistry of a reaction; the nearby group itself does not undergo any evident change during the reaction.

## REDUCTION OF MONOSACCHARIDES

The carbonyl groups of monosaccharides undergo many reactions characteristic of simple ketones and aldehydes. For example, treatment of an aldose or a ketose with  $\text{NaBH}_4$  reduces it to a polyalcohol called an **alditol**. The reduction occurs by interception of the open-chain monosaccharide present in equilibrium between the aldehyde/ketone and the hemiacetal. Although only a small amount of open-chain form is present at any one time, that small amount is reduced, more is produced by opening of the pyranose form, that additional amount is reduced, and so on until the entire sample has undergone reaction.



D-Glucitol, the alditol produced on reduction of D-glucose, is itself a natural product that has been isolated from many fruits and berries. It is used in many foods under the name D-sorbitol as an artificial sweetener and sugar substitute.

PROBLEM .....

- 24.17 How can you account for the fact that reduction of D-galactose (Figure 24.2) with  $\text{NaBH}_4$  leads to an alditol that is optically inactive?

PROBLEM .....

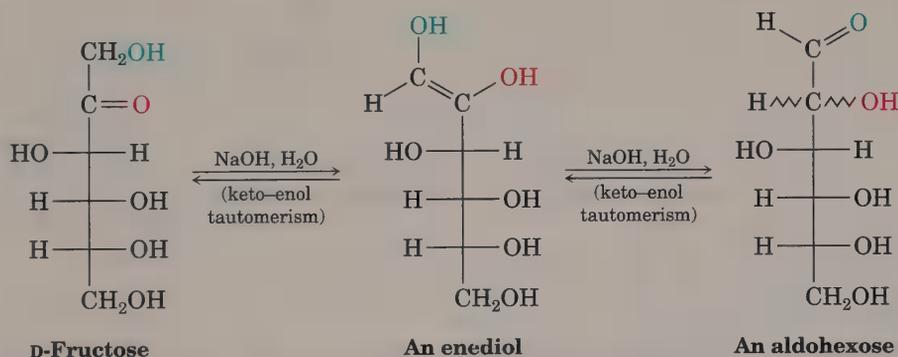
- 24.18 Reduction of L-gulose with  $\text{NaBH}_4$  leads to the same alditol (D-glucitol) as reduction of D-glucose. Explain.
- .....

## OXIDATION OF MONOSACCHARIDES

Like other aldehydes, aldoses are easily oxidized to yield carboxylic acids. Aldoses react with Tollens' reagent ( $\text{Ag}^+$  in aqueous ammonia), Fehling's reagent ( $\text{Cu}^{2+}$  in aqueous sodium tartrate), and Benedict's reagent ( $\text{Cu}^{2+}$  in aqueous sodium citrate) to yield the oxidized sugar and a reduced metallic species. All three reactions serve as simple chemical tests for what are called **reducing sugars** (*reducing* because the sugar reduces the oxidizing agent).

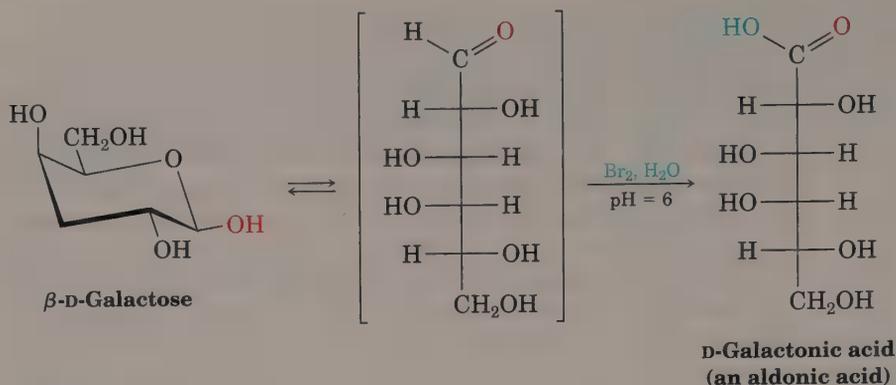
If Tollens' reagent is used, metallic silver is produced as a shiny mirror on the walls of the reaction flask or test tube. If Fehling's or Benedict's reagent is used, a reddish precipitate of cuprous oxide signals a positive result. Some diabetes self-test kits sold in drugstores for home use employ the Benedict test. As little as 0.1% glucose in urine gives a positive test.

All aldoses are reducing sugars because they contain an aldehyde carbonyl group, but some ketoses are reducing sugars as well. For example, fructose reduces Tollens' reagent even though it contains no aldehyde group. This occurs because fructose is readily isomerized to an aldose in basic solution by a series of keto-enol tautomeric shifts (Figure 24.8). Once formed, the aldose is oxidized normally. Glycosides, however, are nonreducing; they don't react with Tollens' reagent because the acetal group can't open to an aldehyde under basic conditions.

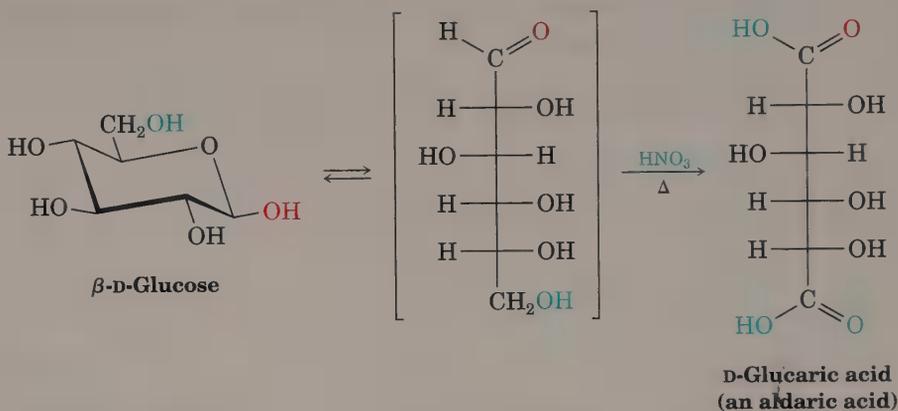


**Figure 24.8** Fructose gives a positive Tollens test as a reducing sugar because it undergoes base-catalyzed keto-enol tautomerism that results in its conversion to an aldohexose.

Although the Tollens and Fehling reactions serve as useful tests for reducing sugars, they don't give good yields of carboxylic acid products because the alkaline conditions used cause decomposition of the carbohydrate skeleton. It has been found, however, that a buffered solution of aqueous bromine oxidizes aldoses to monocarboxylic acids called **aldonic acids**. The reaction is specific for aldoses; ketoses are not oxidized by bromine water.



If a more powerful oxidizing agent such as warm dilute nitric acid is used, aldoses are oxidized to dicarboxylic acids called **aldaric acids**. Both the  $-\text{CHO}$  group at C1 and the terminal  $-\text{CH}_2\text{OH}$  group are oxidized in this reaction.



A summary of the various kinds of carbohydrate derivatives is shown in Figure 24.9.

PROBLEM.....

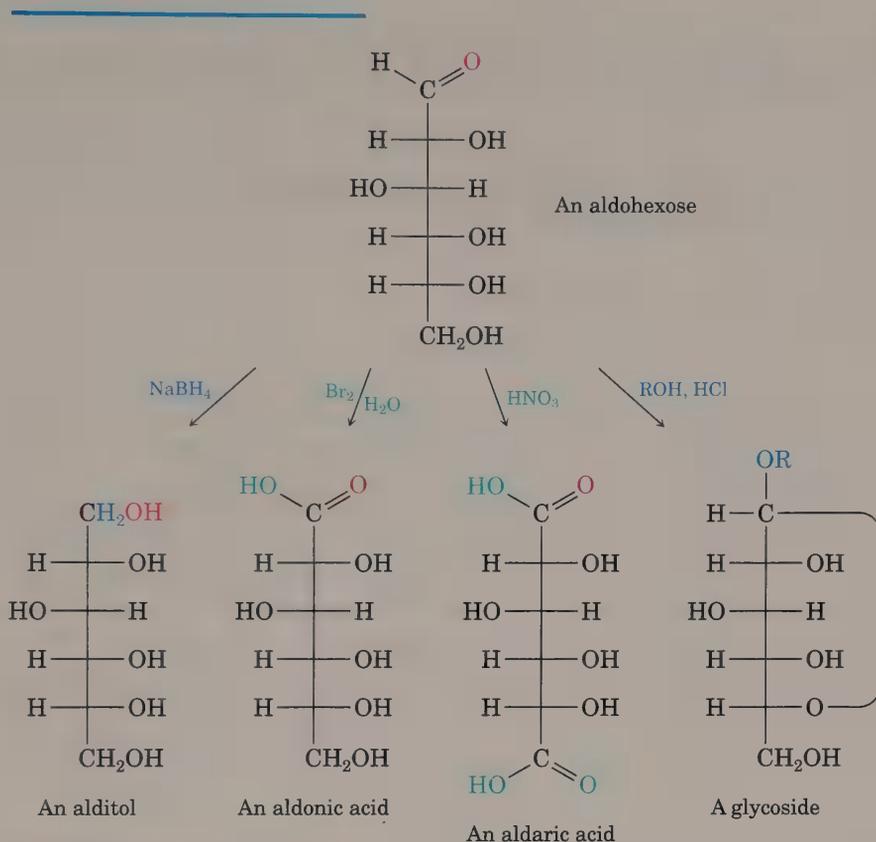
- 24.19 D-Glucose yields an optically active aldaric acid on treatment with nitric acid, but D-allose yields an optically inactive aldaric acid. Explain.

PROBLEM.....

- 24.20 Which of the other six D-aldohexoses yield optically active aldaric acids on oxidation, and which yield meso aldaric acids? (See Problem 24.19.)

## CHAIN LENGTHENING: THE KILIANI-FISCHER SYNTHESIS

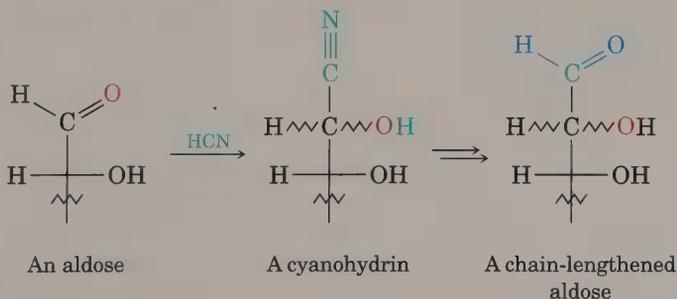
Much early activity in carbohydrate chemistry was devoted to unraveling the various stereochemical relationships among monosaccharides. One of the most important methods used was the **Kiliani-Fischer synthesis**,



**Figure 24.9** A summary of carbohydrate derivatives.

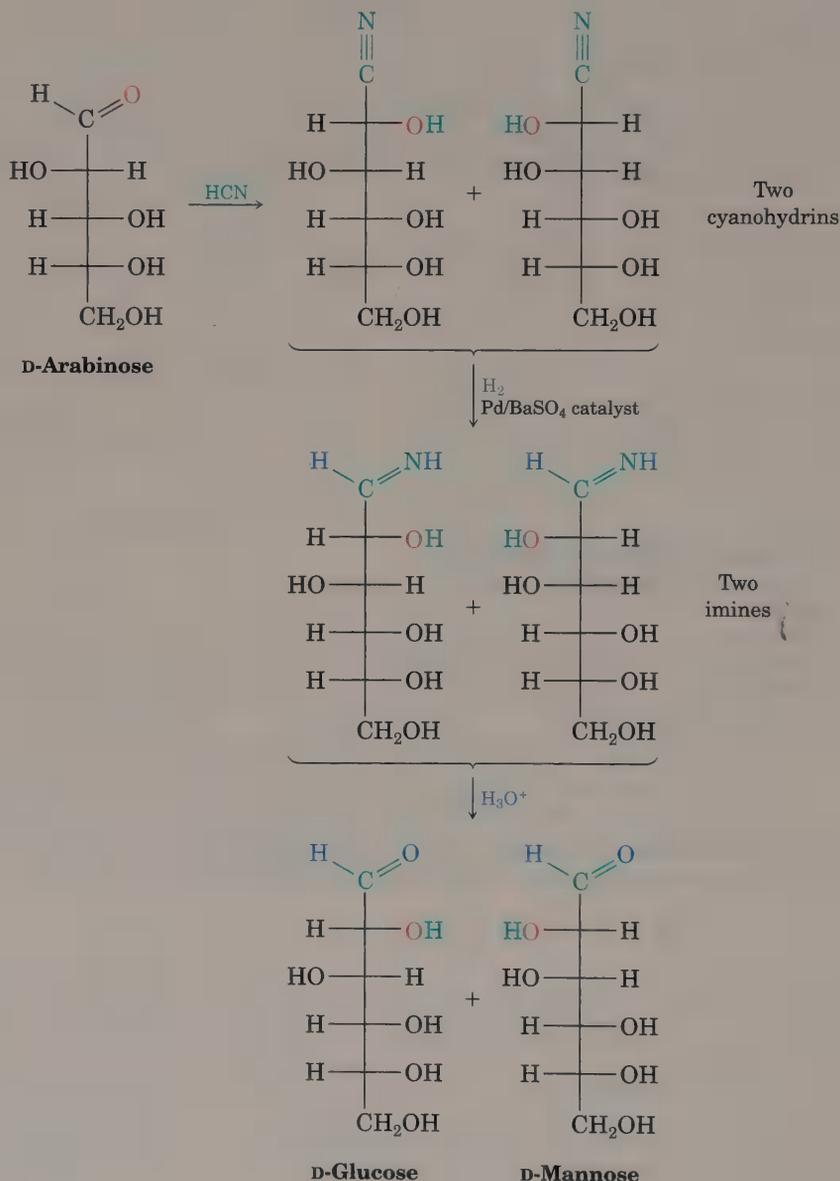
which results in the lengthening of an aldose chain by one carbon atom. For example, an aldopentose is converted by the Kiliani–Fischer synthesis into an aldohexose.

Discovery of the chain-lengthening sequence was initiated by the observation of Heinrich Kiliani<sup>6</sup> in 1886 that aldoses react with HCN to form cyanohydrins (Section 19.9). Emil Fischer immediately realized the importance of Kiliani's discovery and in 1890 published a method for converting the cyanohydrin nitrile group into an aldehyde group.



<sup>6</sup>Heinrich Kiliani (1855–1945); b. Würzburg, Germany; Ph.D. Munich (Erlenmeyer); professor, University of Freiburg.

Conversion of the nitrile into an aldehyde is carried out by catalytic hydrogenation over a palladium catalyst to yield an imine intermediate, followed by hydrolysis. Note that the initial cyanohydrin is a mixture of stereoisomers at the new stereogenic center. Thus, *two* new aldoses, differing only in their stereochemistry at C2, result from Kiliani–Fischer synthesis. For example, chain extension of *D*-arabinose yields a mixture of *D*-glucose and *D*-mannose (Figure 24.10).



**Figure 24.10** Kiliani–Fischer chain lengthening of *D*-arabinose leads to a mixture of *D*-glucose and *D*-mannose.

PROBLEM .....

24.21 What product(s) would you expect from Kiliani–Fischer reaction of D-ribose?

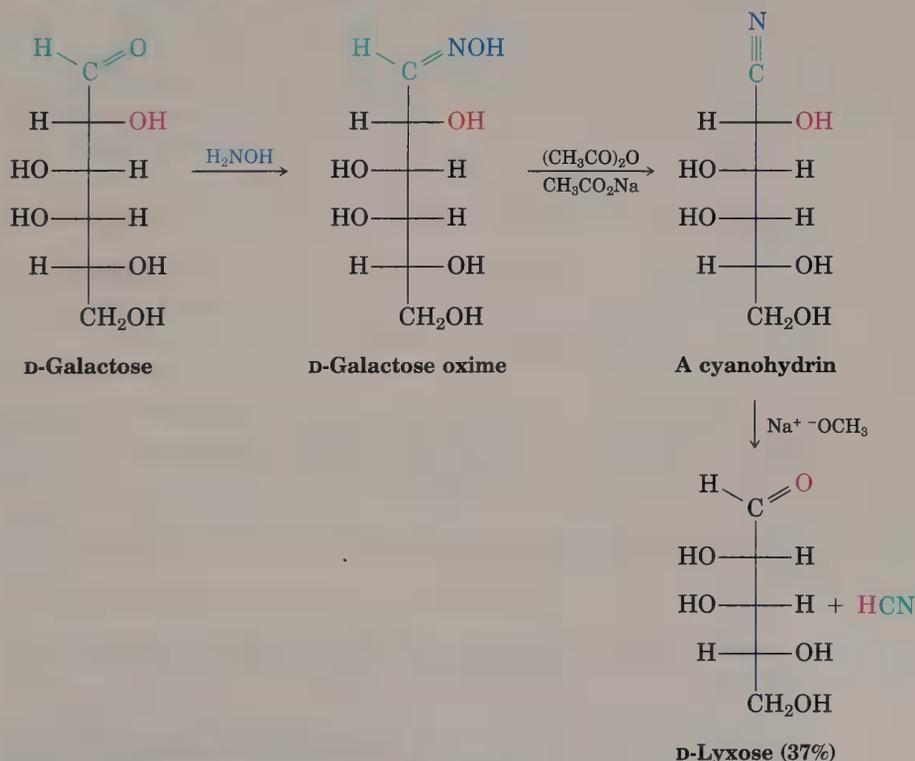
PROBLEM .....

24.22 What aldopentose would give a mixture of L-gulose and L-idose on Kiliani–Fischer chain extension?

## CHAIN SHORTENING: THE WOHL DEGRADATION

Just as the Kiliani–Fischer synthesis lengthens an aldose chain by one carbon, the **Wohl**<sup>7</sup> **degradation** shortens an aldose chain by one carbon. The Wohl degradation is almost exactly the opposite of the Kiliani–Fischer sequence: The aldose aldehyde carbonyl group is first converted into a nitrile group, and the resulting cyanohydrin loses HCN under basic conditions (a retro nucleophilic addition reaction).

Conversion of the aldehyde into a nitrile is accomplished by treatment of an aldose with hydroxylamine, followed by dehydration of the oxime product with acetic anhydride. The Wohl degradation does not give particularly high yields of chain-shortened aldoses, but the reaction is general for all aldopentoses and aldohexoses. For example, D-galactose is converted by Wohl degradation into D-lyxose:



<sup>7</sup>Alfred Wohl (1863–1933); b. Graudenz; Ph.D. Berlin (Hofmann); professor, University of Danzig.

## PROBLEM .....

24.23 What two D-aldopentoses yield D-threose on Wohl degradation?

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## 24.9 Stereochemistry of Glucose: The Fischer Proof

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In the late 1800s, the stereochemical theories of van't Hoff and Le Bel on the tetrahedral geometry of carbon were barely a decade old. Modern chromatographic methods of product purification were unknown, and modern spectroscopic techniques of structure determination were undreamed of. Despite these obstacles, Emil Fischer published in 1891 what stands today as one of the finest examples of chemical logic ever recorded: a structure proof of the stereochemistry of glucose. Let's follow Fischer's logic and see how he arrived at his conclusions.

**1. (+)-Glucose is an aldohexose.** Glucose has four stereogenic centers and can therefore be any one of  $2^4 = 16$  possible stereoisomers (eight pairs of enantiomers). Since no method was available at the time for determining the *absolute* three-dimensional stereochemistry of a molecule, Fischer realized that the best he could do would be to limit his choices for the structure of glucose to a pair of enantiomers. He decided to simplify matters by considering only the eight enantiomers having the C5 hydroxyl group on the right in Fischer projections—what we now call the D series. Fischer was well aware that this arbitrary choice of D-series stereochemistry had only a 50/50 chance of being right, but it was finally shown some 60 years later by the use of sophisticated X-ray techniques that the choice was indeed correct.

The four possible D-aldopentoses and the eight possible D-aldohexoses derived from them by Kiliani–Fischer synthesis are shown in Figure 24.11. One of the eight aldohexoses is glucose, but which one?

**2. Arabinose, an aldopentose, is converted by Kiliani–Fischer chain extension into a mixture of glucose and mannose.** This means that glucose and mannose have the same stereochemistry at C3, C4, and C5, and differ only at C2. Glucose and mannose are therefore represented by one of the pairs of structures 1 and 2, 3 and 4, 5 and 6, or 7 and 8 in Figure 24.11.

**3. Arabinose is converted by treatment with warm nitric acid into an optically active aldaric acid.** Of the four aldopentoses (A, B, C, and D in Figure 24.11), A and C give optically inactive meso aldaric acids when oxidized, whereas B and D give optically active products. Thus, arabinose must be either B or D, and mannose and glucose must therefore be either 3 and 4 or 7 and 8 (Figure 24.12 on page 942).

D-Aldopentoses

D-Aldohexoses

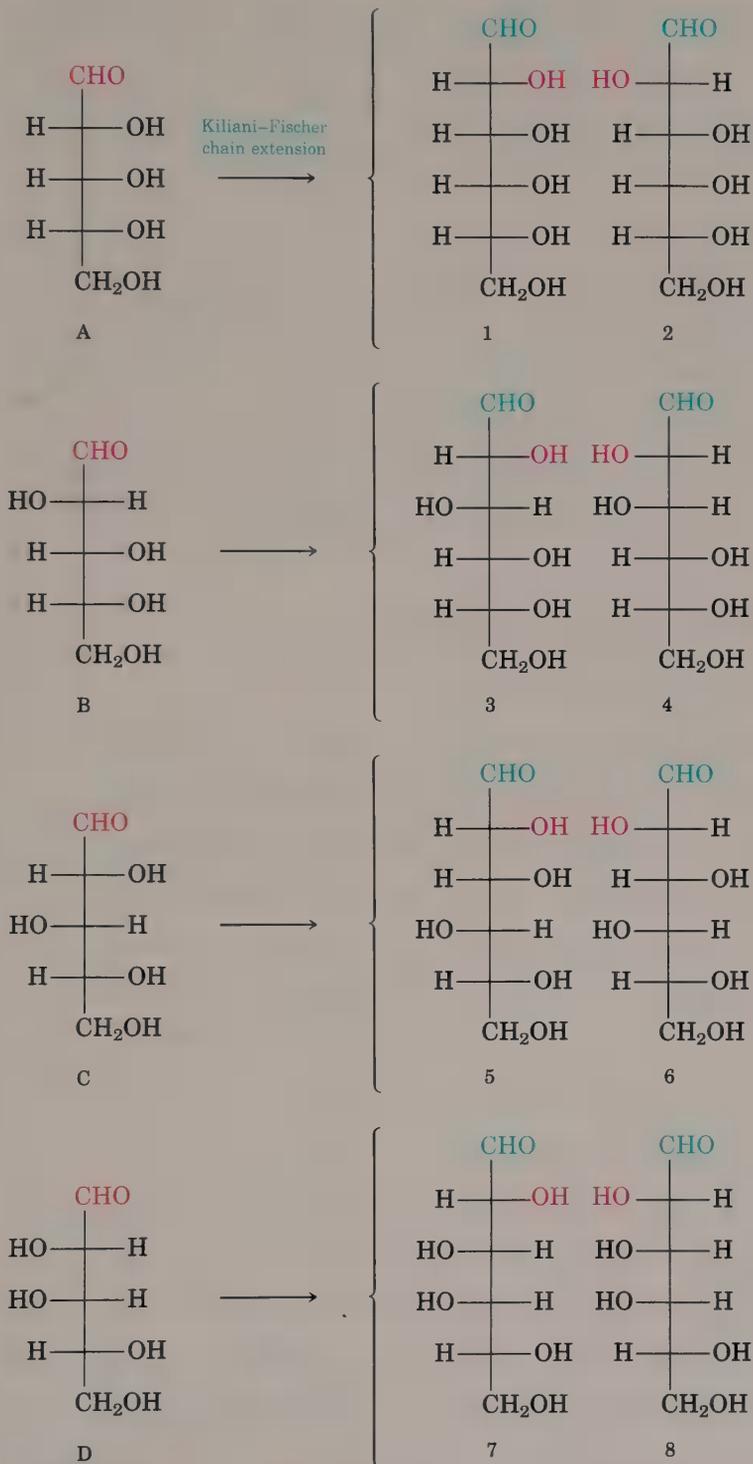
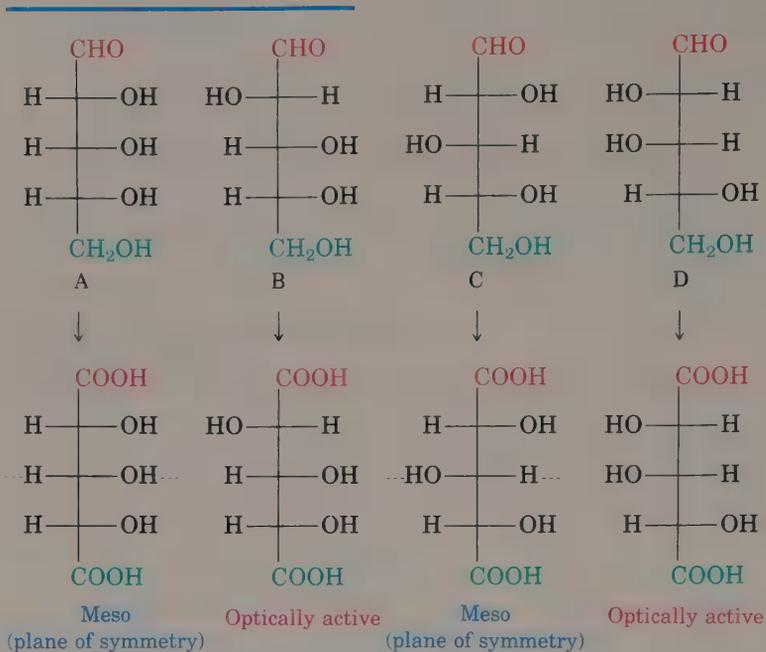
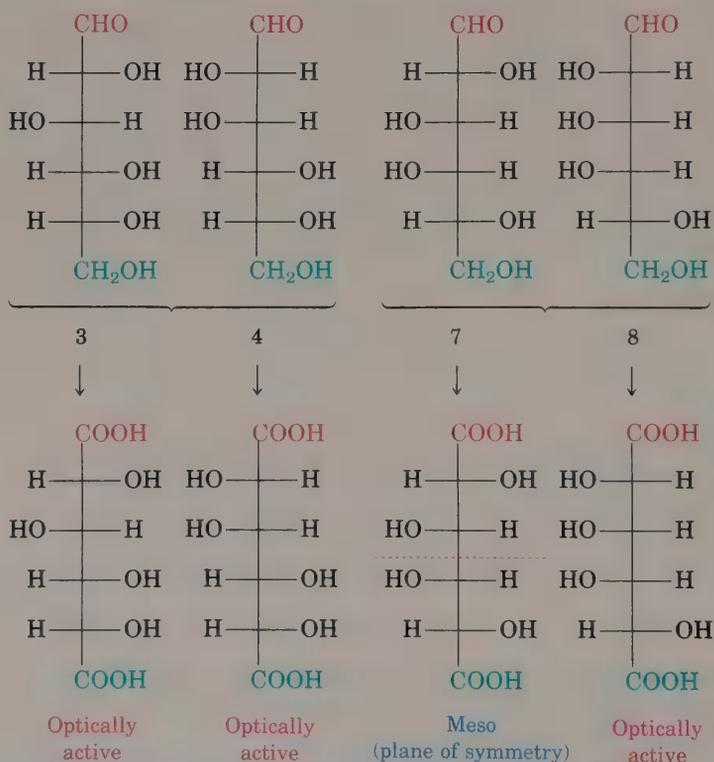


Figure 24.11 The four D-aldopentoses and the eight D-aldohexoses.



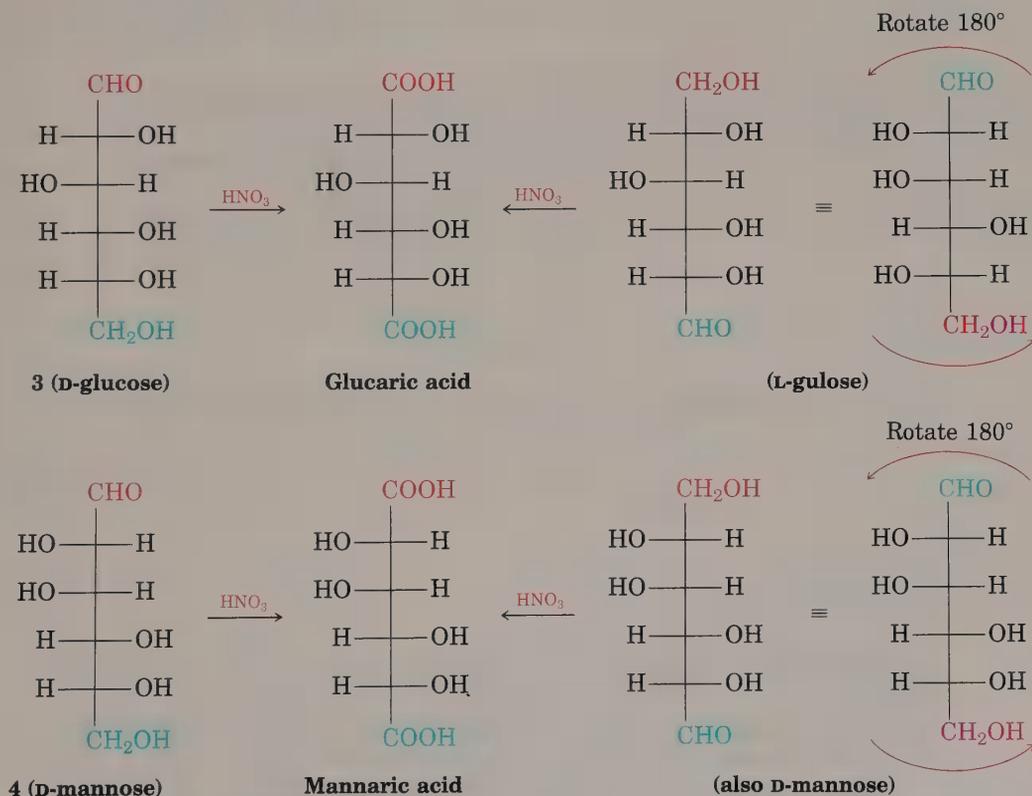
**Figure 24.12** Oxidation of aldopentoses to aldaric acids. Only structures B and D lead to optically active products.



**Figure 24.13** Oxidation of aldohexoses to aldaric acids. Only the pair of structures 3 and 4 both give optically active products.

4. Both glucose and mannose are oxidized by warm nitric acid to optically active aldaric acids. Of the possibilities left at this point, the pair represented by structures 3 and 4 would *both* be oxidized to optically active aldaric acids, but the pair represented by 7 and 8 would not *both* give optically active products. Compound 7 would give an optically inactive meso aldaric acid (Figure 24.13). Thus, glucose and mannose must be 3 and 4, though we can't yet tell which is which.

5. One of the other 15 aldohexoses is converted by nitric acid oxidation to the same aldaric acid as that derived from glucose. How can two different aldohexoses give the same aldaric acid? Since aldaric acids have  $-\text{COOH}$  groups at both ends of the carbon chain, there is no way to tell which was originally the  $-\text{CHO}$  end and which was the  $-\text{CH}_2\text{OH}$  end. Thus, a given aldaric acid can have two different precursors. The aldaric acid from compound 3, for example, might also come from oxidation of a second aldohexose, and the aldaric acid from compound 4 might come from oxidation of a second aldohexose (Figure 24.14):



**Figure 24.14** There is another aldohexose (L-gulose) that can produce the same aldaric acid as compound 3 can, but there is no other aldohexose that can produce the same aldaric acid as compound 4 can. Thus, glucose has structure 3.

If we look carefully at the aldaric acids derived from compounds 3 and 4, we find that the aldaric acid derived from 3 could also come from oxidation of another aldohexose (L-gulose), but that the aldaric acid derived from 4 could not. The “other” aldohexose that could produce the same aldaric acid as that from compound 4 is in fact identical to 4 by a simple  $180^\circ$  rotation. Thus, glucose must have structure 3 and mannose must have structure 4 (Figure 24.14).

Reasoning similar to that described for glucose allowed Fischer to determine the stereochemistry of 12 of the 16 aldohexoses. For this remarkable achievement, he was awarded the 1902 Nobel Prize in chemistry.

PROBLEM .....

- 24.24 The structures of the four aldopentoses, A, B, C, and D, are shown in Figure 24.11. In light of point 2 presented by Fischer, what is the structure of arabinose? In light of point 3, what is the structure of lyxose, another aldopentose that yields an optically active aldaric acid?

PROBLEM .....

- 24.25 The aldotetrose D-erythrose yields a mixture of D-ribose and D-arabinose on Kiliani–Fischer chain extension.
- What is the structure of D-ribose?
  - What is the structure of D-xylose, the fourth possible aldopentose?
  - What is the structure of D-erythrose?
  - What is the structure of D-threose, the other possible aldotetrose?
- .....

## 24.10 Disaccharides

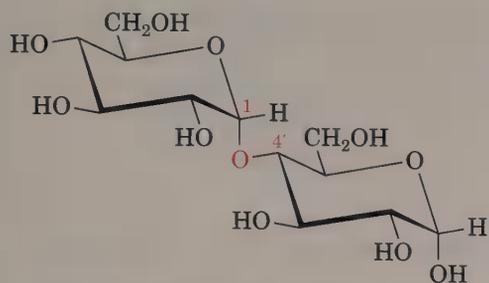
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We saw earlier that reaction of a monosaccharide hemiacetal with an alcohol yields a glycoside in which the anomeric hydroxyl group is replaced by an alkoxy substituent. If the alcohol is itself a sugar, however, the glycosidic product is a disaccharide.

### CELLOBIOSE AND MALTOSE

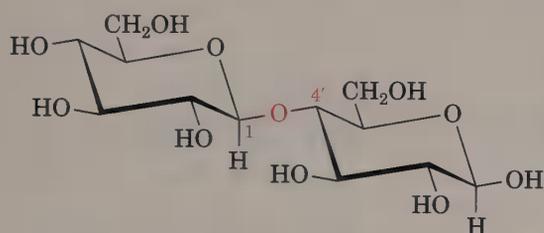
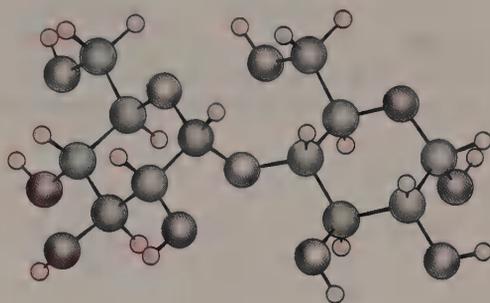
Disaccharides are compounds that contain a glycosidic acetal bond between C1 of one sugar and a hydroxyl group at *any* position on the other sugar. A glycosidic bond between C1 of the first sugar and C4 of the second sugar is particularly common but is by no means required. Such a bond is called a **1,4' link**, read as “one, four-prime.” The prime superscript indicates that the 4' position is on a different sugar than the nonprime 1 position is.

A glycosidic bond to the anomeric carbon can be either alpha or beta. Maltose, the disaccharide obtained by enzyme-catalyzed hydrolysis of starch, consists of two D-glucopyranoses joined by a 1,4'- $\alpha$ -glycoside bond. Cellobiose, the disaccharide obtained by partial hydrolysis of cellulose, consists of two D-glucopyranoses joined by a 1,4'- $\beta$ -glycoside bond.



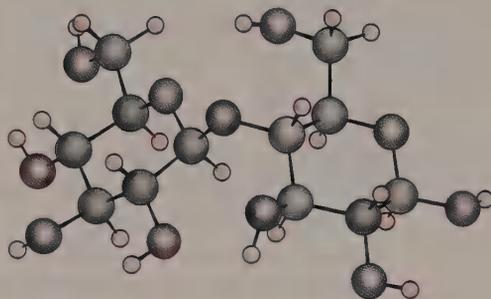
**Maltose, a 1,4'- $\alpha$ -glycoside**

[4-*O*-( $\alpha$ -D-glucopyranosyl)- $\alpha$ -D-glucopyranose]

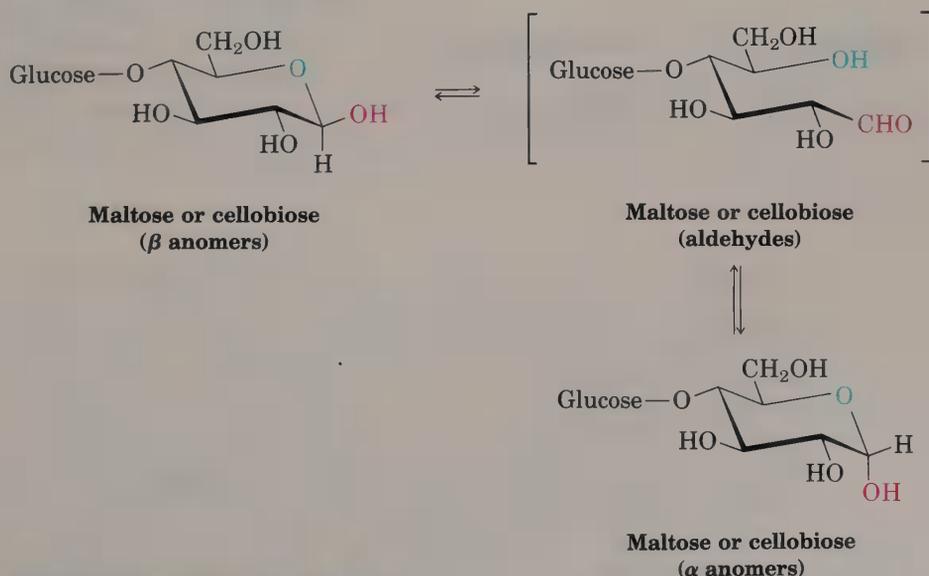


**Cellobiose, a 1,4'- $\beta$ -glycoside**

[4-*O*-( $\beta$ -D-glucopyranosyl)- $\beta$ -D-glucopyranose]



Maltose and cellobiose are both reducing sugars because the anomeric carbons on the right-hand glucopyranose subunits have hemiacetal groups. Both are therefore in equilibrium with aldehyde forms, which can reduce Tollens' or Fehling's reagent. For a similar reason, both maltose and cellobiose exhibit mutarotation of alpha and beta anomers of the glucopyranose unit on the right (Figure 24.15).

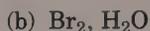
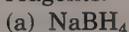


**Figure 24.15** Mutarotation of maltose and cellobiose.

Despite the similarities of their structures, cellobiose and maltose are dramatically different biologically. Cellobiose can't be digested by humans and can't be fermented by yeast. Maltose, however, is digested without difficulty and is fermented readily.

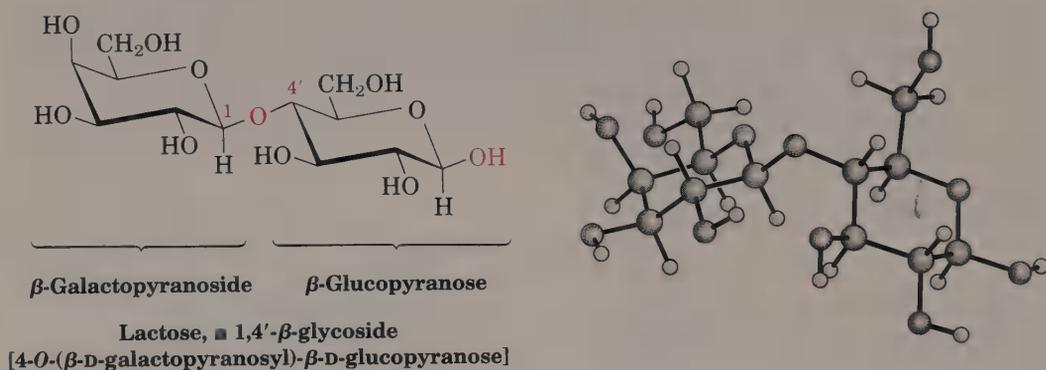
## PROBLEM

24.26 Show the product you would obtain from the reaction of cellobiose with these reagents:



## LACTOSE

Lactose is a disaccharide that occurs naturally in both human and cow's milk. It is widely used in baking and in commercial infant-milk formulas. Like cellobiose and maltose, lactose is a reducing sugar. It exhibits mutarotation and is a 1,4'- $\beta$ -linked glycoside. Unlike cellobiose and maltose, however, lactose contains two *different* monosaccharide units. Acidic hydrolysis of lactose yields one equivalent of D-glucose and one equivalent of D-galactose; the two are joined by a  $\beta$ -glycosidic bond between C1 of galactose and C4 of glucose.

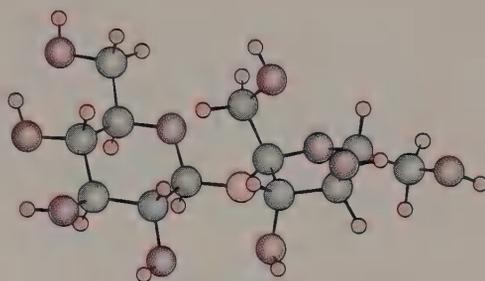
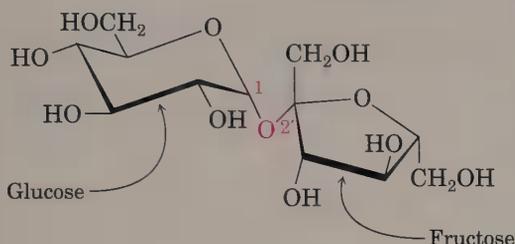


## SUCROSE

Sucrose, ordinary table sugar, is probably the single most abundant pure organic chemical in the world and the one most widely known to nonchemists. Whether from sugar cane (20% by weight) or sugar beets (15% by weight), and whether raw or refined, all table sugar is sucrose.

Sucrose is a disaccharide that yields one equivalent of glucose and one equivalent of fructose on acidic hydrolysis. This 1:1 mixture of glucose and fructose is often referred to as *invert sugar*, because the sign of optical rotation changes (inverts) during the hydrolysis from sucrose ( $[\alpha]_D = +66.5^\circ$ ) to a glucose/fructose mixture ( $[\alpha]_D \approx -22.0^\circ$ ). Certain insects such as honeybees have enzymes called *invertases* that catalyze the hydrolysis of sucrose to a glucose + fructose mixture. Honey, in fact, is primarily a mixture of glucose, fructose, and sucrose.

Unlike most other disaccharides, sucrose is not a reducing sugar and does not exhibit mutarotation. These observations imply that sucrose has no hemiacetal groups and suggest that glucose and fructose must both be glycosides. This can happen only if the two sugars are joined by a glycoside link between C1 of glucose and C2 of fructose.



Sucrose, a 1,2'-glycoside

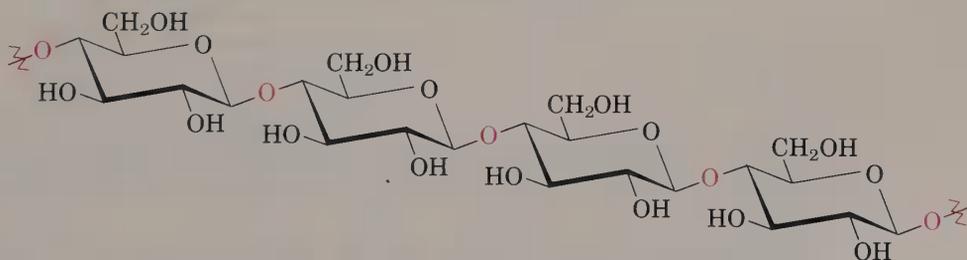
[2-*O*-( $\alpha$ -D-glucopyranosyl)- $\beta$ -D-fructofuranoside]

## 24.11 Polysaccharides

Polysaccharides are carbohydrates in which tens, hundreds, or even thousands of simple sugars are linked together through glycoside bonds. Since they have no free anomeric hydroxyls (except for one at the end of the chain), polysaccharides aren't reducing sugars and don't show mutarotation. Cellulose and starch are the two most widely occurring polysaccharides.

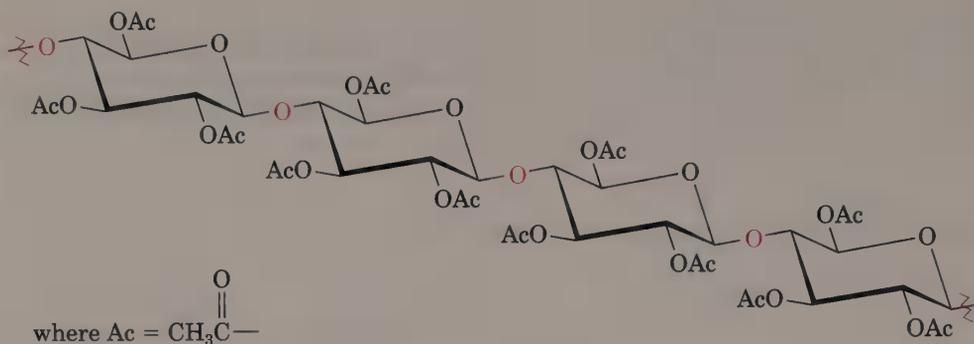
### CELLULOSE

Cellulose consists simply of D-glucose units linked by the 1,4'- $\beta$ -glycoside bonds we saw in cellobiose. Several thousand glucose units are linked to form one large molecule, and different molecules can then interact to form a large aggregate structure held together by hydrogen bonds.



Cellulose, ■ 1,4'-*O*-( $\beta$ -D-glucopyranoside) polymer

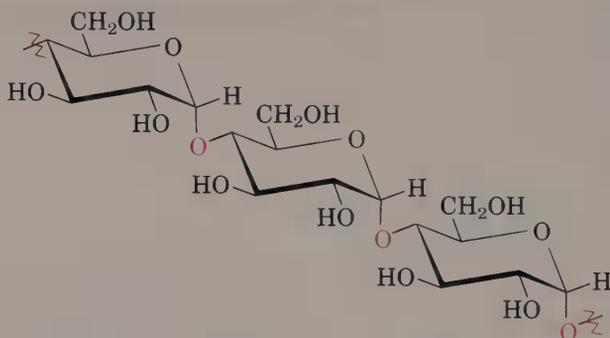
Nature uses cellulose primarily as a structural material to impart strength and rigidity to plants. Wood, leaves, grasses, and cotton are primarily cellulose. Cellulose also serves as raw material for the manufacture of cellulose acetate, known commercially as acetate rayon.



A segment of cellulose acetate (acetate rayon)

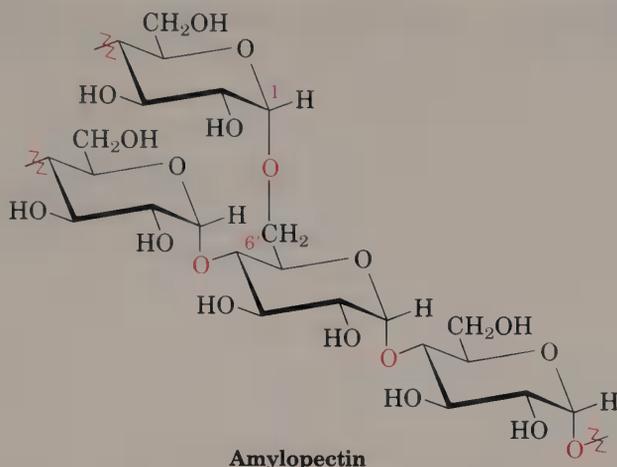
## STARCH AND GLYCOGEN

Starch is a polymer of glucose in which monosaccharide units are linked by the 1,4'- $\alpha$ -glycoside bonds we saw in maltose. Starch can be separated into two fractions: a fraction soluble in cold water, called *amylopectin*, and a fraction insoluble in cold water, called *amylose*. Amylose, which accounts for about 20% by weight of starch, consists of several hundred glucose molecules linked together by 1,4'- $\alpha$ -glycoside bonds.

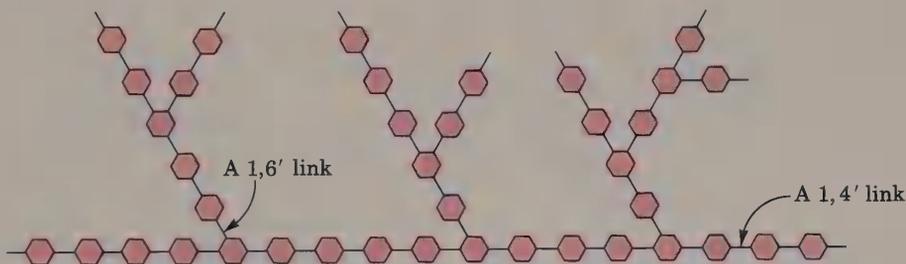
Amylose, ■ 1,4'-O-( $\alpha$ -D-glucopyranoside) polymer

Amylopectin, which accounts for the remaining 80% of starch, is more complex in structure than amylose. Unlike cellulose and amylose, which are linear polymers, amylopectin contains 1,6'- $\alpha$ -glycoside *branches* approximately every 25 glucose units. As a result, amylopectin has an exceedingly complex three-dimensional structure. Nature uses starch as the means by which plants store energy for later use. Potatoes, corn, and cereal grains contain large amounts of starch.

When eaten, starch is digested in the mouth and stomach by enzymes called *glycosidases*, which catalyze the hydrolysis of glycoside bonds and release individual molecules of glucose. Like most enzymes, glycosidases are highly selective in their action. They hydrolyze only the  $\alpha$ -glycoside links in starch and leave the  $\beta$ -glycoside links in cellulose untouched. Thus, humans can eat potatoes and grains but not grass and wood.



Glycogen is a polysaccharide that serves the same energy-storage purpose in animals that starch serves in plants. Dietary carbohydrate that isn't needed for immediate energy is converted by the body to glycogen for long-term storage. Like the amylopectin found in starch, glycogen contains a complex three-dimensional structure with both 1,4' and 1,6' links (Figure 24.16). Glycogen molecules are larger than those of amylopectin—up to 100,000 glucose units—and contain even more branches.



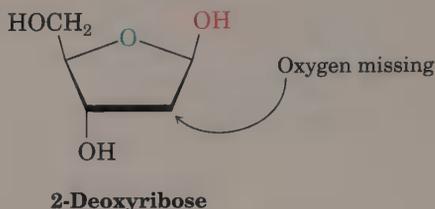
**Figure 24.16** A representation of the structure of glycogen. The hexagons represent glucose units linked by 1,4' and 1,6' acetal bonds.

## 24.12 Other Important Carbohydrates

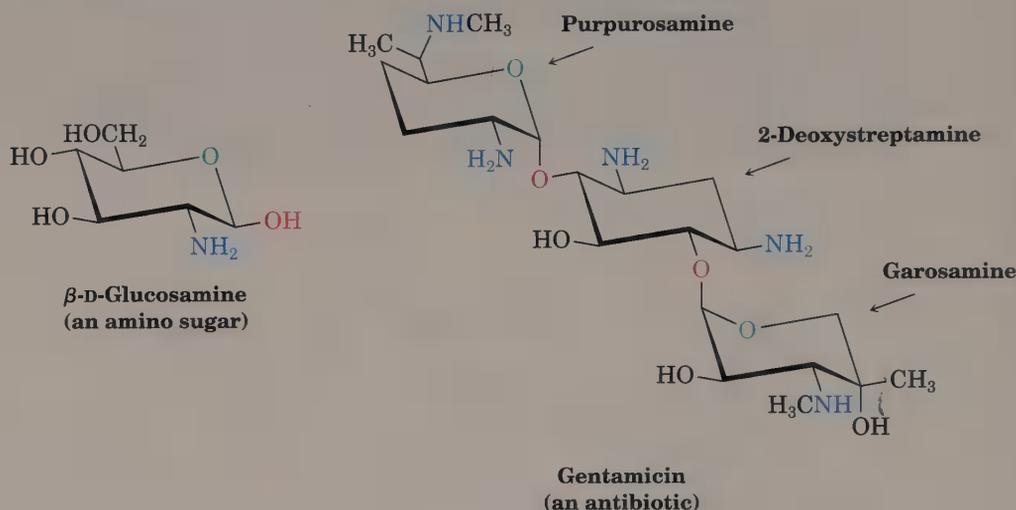
In addition to the common carbohydrates mentioned in previous sections, there are a variety of important carbohydrate-derived materials whose structures have been chemically modified. Their structural resemblance to sugars is clear, but they aren't simple aldoses or ketoses.

**Deoxy sugars** differ from normal sugars by having one of their oxygen atoms “missing.” In other words, an  $-OH$  group is replaced by an  $-H$ .

2-Deoxyribose, a sugar found in DNA (deoxyribonucleic acid), is the most important deoxy sugar. Note that 2-deoxyribose adopts a furanose (five-membered) form.



**Amino sugars** such as D-glucosamine have one of their  $-OH$  groups replaced by an  $-NH_2$ . The *N*-acetyl amide derived from D-glucosamine is the monosaccharide unit from which *chitin*, the hard crust that protects insects and shellfish, is built. Still other amino sugars are found in antibiotics such as streptomycin and gentamicin.



## 24.13 Carbohydrates on Cell Surfaces

For many years, carbohydrates were thought to be dull compounds whose only biological purposes were to serve as structural materials and as energy sources. Although carbohydrates do indeed fill these two roles, it's now known that they perform many other important biochemical functions as well. For example, polysaccharides are centrally involved in the critical process by which one cell type recognizes another. Small polysaccharide chains, covalently bound by glycosidic links to hydroxyl groups on proteins (*glycoproteins*), act as biochemical labels on cell surfaces, as illustrated by the human blood-group antigens.

It has been known for over 90 years that human blood can be classified into four blood-group types, A, B, AB, and O, and that blood from a donor of one type can't be transfused into a recipient having another type unless

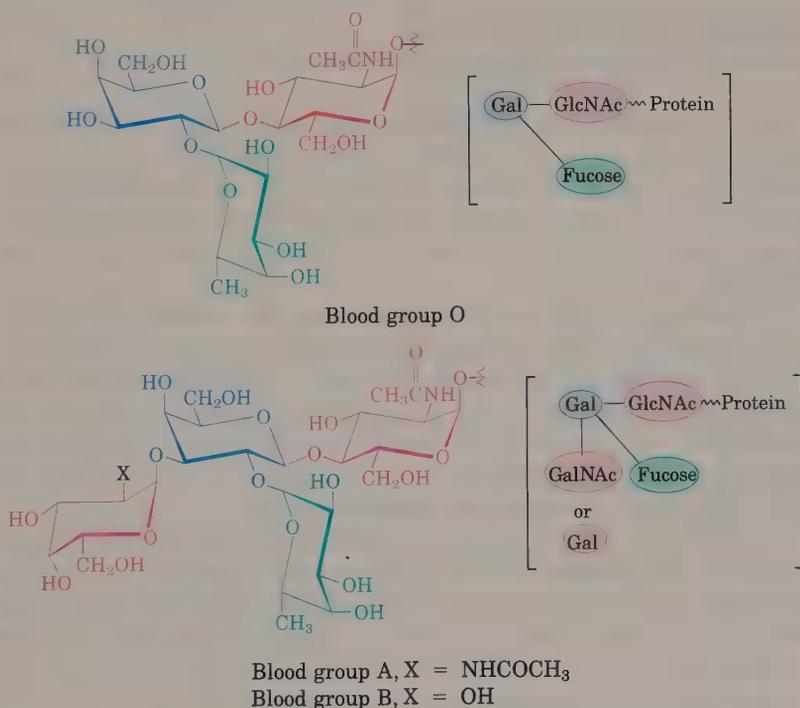
the two types are compatible (Table 24.1). Should an incompatible mix be made, the red blood cells clump together, or *agglutinate*.

**TABLE 24.1** Human Blood-Group Compatibilities

Donor blood type	Acceptor blood type			
	A	B	AB	O
A	○	×	○	×
B	×	○	○	×
AB	×	×	○	×
O	○	○	○	○

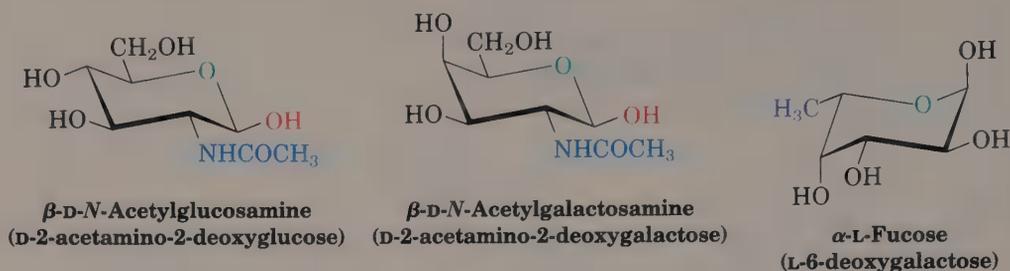
○ = compatible; × = incompatible

This agglutination of incompatible types of red blood cells, which indicates that the body's immune system has recognized the presence of foreign substances in the body and has formed antibodies against them, results from the presence of polysaccharide markers on the surface of the cells. Types A, B, and O red blood cells each have characteristic markers, called *antigenic determinants*; type AB cells have both type A and type B markers. The structures of all three blood-group determinants are shown in Figure 24.17.



**Figure 24.17** Structures of the A, B, and O blood-group antigenic determinants (Gal = D-galactose; GlcNAc = N-acetylglucosamine; GalNAc = N-acetylgalactosamine).

Note that some unusual carbohydrates are involved. Thus, all three blood-group antigenic determinants contain *N*-acetylamino sugars as well as the unusual monosaccharide L-fucose.



The antigenic determinant of blood group O is a trisaccharide, whereas the determinants of blood groups A and B have an additional sugar attached at C3 of the galactose unit. The type A and B determinants differ only in the substitution of an acetylamino group ( $-\text{NHCOCH}_3$ ) for a hydroxyl in the terminal galactose residue.

Elucidation of the role of carbohydrates in cell recognition is an exciting area of current research that offers hope of breakthroughs in the understanding of a wide range of diseases from bacterial infections to cancer. The potential benefits of work in this field are enormous.

## 24.14 Summary and Key Words

**Carbohydrates** are polyhydroxy aldehydes and ketones. They can be classified according to the number of carbon atoms and the kind of carbonyl group they contain. Glucose, for example, is an **aldohexose**, a six-carbon aldehyde sugar. Monosaccharides are further classified as either **D** or **L sugars**, depending on the stereochemistry of the chiral carbon atom farthest from the carbonyl group.

Monosaccharides normally exist as **cyclic hemiacetals** rather than as open-chain aldehydes or ketones. The hemiacetal linkage results from reaction of the carbonyl group with a hydroxyl group three or four carbon atoms away. A five-membered-ring hemiacetal is called a **furanose**, and a six-membered-ring hemiacetal is called a **pyranose**. Cyclization leads to the formation of a new stereogenic center and production of two diastereomeric hemiacetals called **alpha** and **beta anomers**.

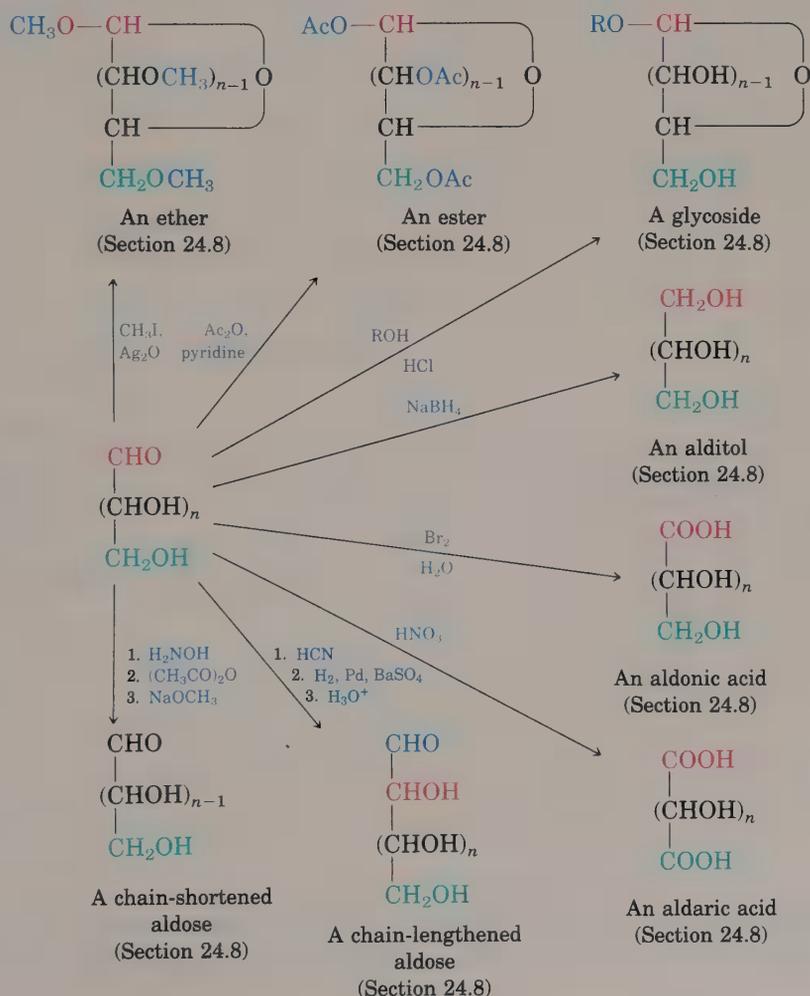
Stereochemical relationships among monosaccharides are portrayed in several ways. **Fischer projections** display chiral carbon atoms as a pair of crossed lines; cyclic **Haworth projections** provide a more accurate view. Any group to the right in a Fischer projection is down in a Haworth projection, and any group to the left in a Fischer projection is up in a Haworth projection.

Much of the chemistry of monosaccharides is the now-familiar chemistry of alcohols and aldehydes/ketones. Thus, the hydroxyl groups of carbohydrates form esters and ethers in the normal way. The carbonyl group of a

monosaccharide can be reduced with sodium borohydride to form an **alditol**, oxidized with bromine water to form an **aldonic acid**, oxidized with warm nitric acid to form an **aldaric acid**, or treated with an alcohol in the presence of acid to form a **glycoside**. Monosaccharides can also be chain-lengthened by the multistep **Kiliani–Fischer synthesis** and can be chain-shortened by the **Wohl degradation**.

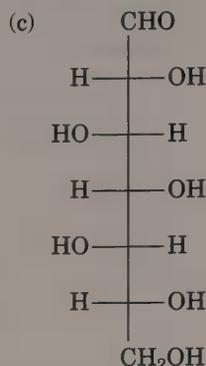
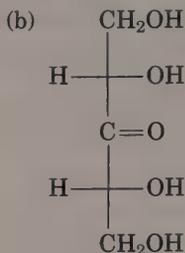
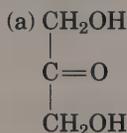
**Disaccharides** are complex carbohydrates in which two simple sugars are linked by a glycoside bond between the anomeric carbon of one unit and a hydroxyl of the second unit. The two sugars can be the same, as in maltose and cellobiose, or different, as in lactose and sucrose. The glycosidic bond can be either  $\alpha$  (maltose) or  $\beta$  (cellobiose, lactose) and can involve any hydroxyl of the second sugar. A 1,4' link is most common (cellobiose, maltose), but others such as 1,2' (sucrose) are also known.

## 24.15 Summary of Reactions



## ADDITIONAL PROBLEMS .....

24.27 Classify the following sugars by type; for example, glucose is an aldohexose.

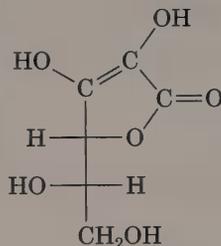


24.28 Write open-chain structures for a ketotetrose and a ketopentose.

24.29 Write an open-chain structure for a deoxyaldohexose.

24.30 Write an open-chain structure for a five-carbon amino sugar.

24.31 Does ascorbic acid (vitamin C) have a D or an L configuration?



Ascorbic acid

24.32 Draw a Haworth projection of ascorbic acid (Problem 24.31).

24.33 Define the following terms, and give an example of each.

(a) Monosaccharide

(b) Anomeric center

(c) Haworth projection

(d) Fischer projection

(e) Glycoside

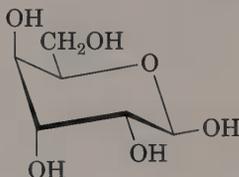
(f) Reducing sugar

(g) Pyranose form

(h) 1,4' Link

(i) D-Sugar

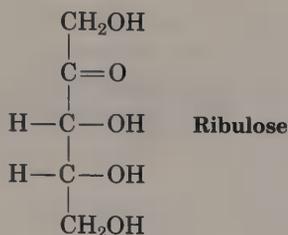
24.34 The following cyclic structure is that of gulose. Is this a furanose or pyranose form? Is it an  $\alpha$  or a  $\beta$  anomer? Is it a D-sugar or an L-sugar?



Gulose

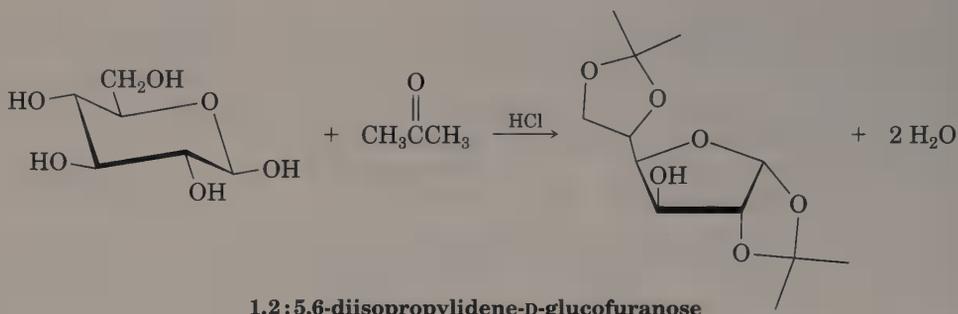
24.35 Uncoil gulose (Problem 24.34), and write it in its open-chain form.

24.36 Draw D-ribulose in its five-membered cyclic  $\beta$ -hemiacetal form.



- 24.37** Look up the structure of D-talose in Figure 24.2, and draw the  $\beta$  anomer in its pyranose form. Identify the ring substituents as axial or equatorial.
- 24.38** Draw structures for the products you would expect to obtain from reaction of  $\beta$ -D-talopyranose with each of the following reagents:
- |  |   |   |
|--|---|---|
| (a) $\text{NaBH}_4$ in $\text{H}_2\text{O}$          | (b) Warm dilute $\text{HNO}_3$                    | (c) $\text{Br}_2$ , $\text{H}_2\text{O}$          |
| (d) $\text{CH}_3\text{CH}_2\text{OH}$ , $\text{HCl}$ | (e) $\text{CH}_3\text{I}$ , $\text{Ag}_2\text{O}$ | (f) $(\text{CH}_3\text{CO})_2\text{O}$ , pyridine |
- 24.39** How many D-2-ketohexoses are possible? Draw them.
- 24.40** One of the D-2-ketohexoses (Problem 24.39) is called *sorbose*. On treatment with  $\text{NaBH}_4$ , sorbose yields a mixture of gulitol and iditol. What is the structure of sorbose?
- 24.41** Another D-2-ketohexose, *psicose*, yields a mixture of allitol and altritol when reduced with  $\text{NaBH}_4$ . What is the structure of psicose?
- 24.42** Fischer prepared the L-gulose needed for his structure proof of glucose in the following way. D-Glucose was oxidized to D-glucaric acid, which can form two six-membered-ring lactones. These were separated and reduced with sodium amalgam to give D-glucose and L-gulose. What are the structures of the two lactones, and which one is reduced to L-gulose?
- 24.43** What other D-aldohexose gives the same alditol as D-talose gives?
- 24.44** Which of the eight D-aldohexoses give the same aldaric acids as their L enantiomers give?
- 24.45** Which of the other three D-aldopentoses gives the same aldaric acid as D-lyxose gives?
- 24.46** Gentiobiose, a rare disaccharide found in saffron and gentian, is a reducing sugar and forms only D-glucose on hydrolysis with aqueous acid. Reaction of gentiobiose with iodomethane and silver iodide yields an octamethyl derivative, which can be hydrolyzed with aqueous acid to give 1 equiv of 2,3,4,6-tetra-O-methyl-D-glucopyranose and 1 equiv of 2,3,4-tri-O-methyl-D-glucopyranose. If gentiobiose contains a  $\beta$ -glycoside link, what is its structure?
- 24.47** Amygdalin, or Laetrile, is a glycoside isolated in 1830 from almond and apricot seeds. It is known as a cyanogenic glycoside since acidic hydrolysis liberates HCN, along with benzaldehyde and two equivalents of D-glucose. Structural studies have shown amygdalin to be a  $\beta$ -glycoside of benzaldehyde cyanohydrin with gentiobiose (Problem 24.46). Draw the structure of amygdalin.
- 24.48** Trehalose is a nonreducing disaccharide that is hydrolyzed by aqueous acid to yield two equivalents of D-glucose. Methylation followed by acidic hydrolysis yields two equivalents of 2,3,4,6-tetra-O-methylglucose. How many possible structures are there for trehalose?
- 24.49** Trehalose (Problem 24.48) is cleaved by enzymes that hydrolyze  $\alpha$ -glycosides but not by enzymes that hydrolyze  $\beta$ -glycosides. What are the structure and the systematic name of trehalose?

- 24.50** Isotrehalose and neotrehalose are chemically similar to trehalose (Problem 24.48) except for the fact that neotrehalose is hydrolyzed only by  $\beta$ -glycosidase enzymes, whereas isotrehalose is hydrolyzed by both  $\alpha$ - and  $\beta$ -glycosidase enzymes. What are the structures of isotrehalose and neotrehalose?
- 24.51** Propose a scheme for the synthesis of gentiobiose methyl glycoside (Problem 24.46), starting from  $\beta$ -D-glucose and methyl 2,3,4-tri-O-acetyl- $\beta$ -D-glucopyranoside.
- 24.52** D-Glucose reacts with acetone in the presence of acid to yield the nonreducing 1,2:5,6-diisopropylidene-D-glucofuranose. Propose a mechanism for this reaction.

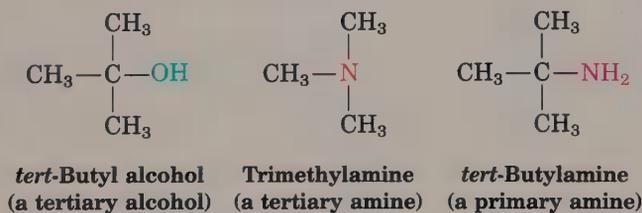


- 24.53** D-Mannose reacts with acetone to give a diisopropylidene derivative that is still reducing toward Tollens' reagent. Propose a likely structure for this derivative.
- 24.54** Propose a mechanism to account for the fact that D-gluconic acid and D-mannonic acid are interconverted when either is heated in pyridine solvent.
- 24.55** The cyclitols are a group of carbocyclic sugar derivatives having the general formulation 1,2,3,4,5,6-cyclohexanehexol. How many stereoisomeric cyclitols are possible? Draw them in Haworth projection.
- 24.56** Compound A is a D-aldopentose that can be oxidized to an optically inactive aldaric acid, B. On Kiliani-Fischer chain extension, A is converted into C and D; C can be oxidized to an optically active aldaric acid E, but D is oxidized to an optically inactive aldaric acid, F. What are the structures of A-F?

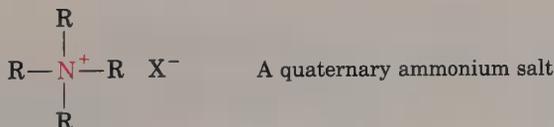
# Chapter 25

## Aliphatic Amines

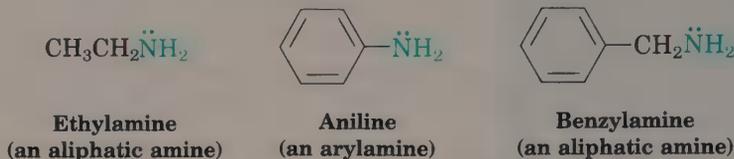
**Amines** are organic derivatives of ammonia in the same way that alcohols and ethers are organic derivatives of water. Amines are classified as **primary** ( $\text{RNH}_2$ ), **secondary** ( $\text{R}_2\text{NH}$ ), or **tertiary** ( $\text{R}_3\text{N}$ ), depending on the number of organic substituents attached to nitrogen. For example, methylamine ( $\text{CH}_3\text{NH}_2$ ) is a primary amine, dimethylamine [ $(\text{CH}_3)_2\text{NH}$ ] is a secondary amine, and trimethylamine [ $(\text{CH}_3)_3\text{N}$ ] is a tertiary amine. Note that this usage of the terms *primary*, *secondary*, and *tertiary* is different from our previous usage. When we speak of a tertiary alcohol or alkyl halide, we refer to the degree of substitution at the alkyl carbon atom; when we speak of a tertiary amine, however, we refer to the degree of substitution at the nitrogen atom.



Compounds with four groups attached to nitrogen are also known, but the nitrogen atom must carry a positive charge. Such compounds are called **quaternary ammonium salts**.

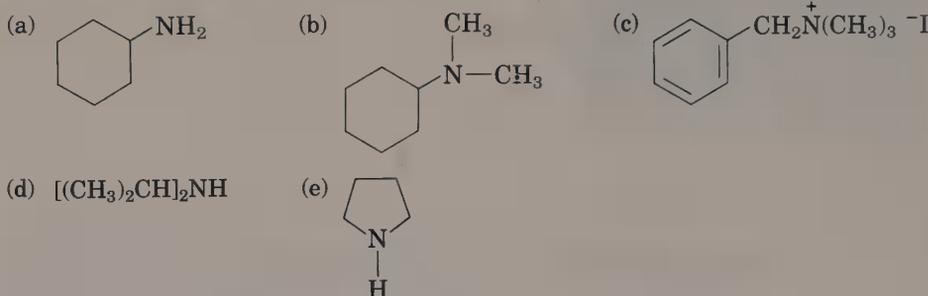


Amines can be either alkyl substituted or aryl substituted. Much of the chemistry of the two classes is similar, but there are sufficient differences that we'll consider them separately. Arylamines will be discussed in Chapter 26.



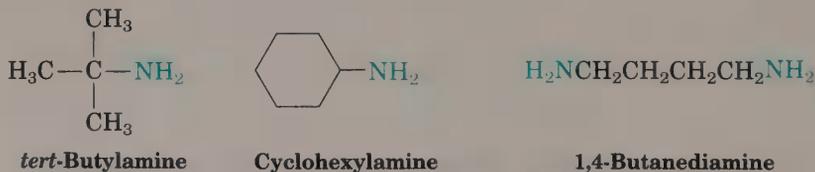
## PROBLEM

25.1 Classify these compounds as either primary, secondary, or tertiary amines, or as quaternary ammonium salts.



## 25.1 Nomenclature of Amines

Primary amines,  $\text{RNH}_2$ , are named in the IUPAC system in several ways, depending on their structure. For simple amines, the suffix *-amine* is added to the name of the alkyl substituent.



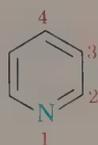
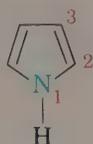
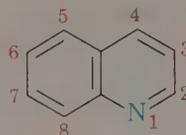
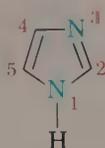
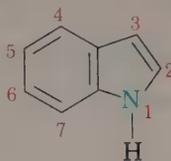
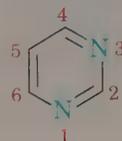
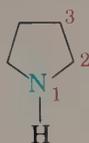
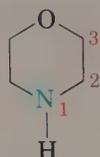
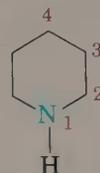
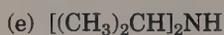
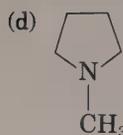
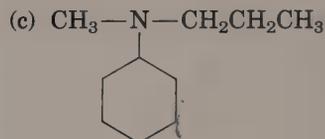
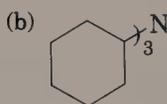
Alternatively, the suffix *-amine* can be used in place of the final *-e* in the name of the parent compound.



4,4-Dimethylcyclohexylamine

Amines with more than one functional group are named by considering the  $-\text{NH}_2$  as an *amino* substituent on the parent molecule.



**Pyridine****Pyrrole****Quinoline****Imidazole****Indole****Pyrimidine****Pyrrolidine****Morpholine****Piperidine****PROBLEM** .....**25.2** Name these compounds by IUPAC rules:**PROBLEM** .....**25.3** Draw structures corresponding to these IUPAC names:

(a) Triethylamine

(b) Triallylamine

(c) *N*-Methylaniline(d) *N*-Ethyl-*N*-methylcyclopentylamine(e) *N*-Isopropylcyclohexylamine(f) *N*-Ethylpyrrole**PROBLEM** .....**25.4** Draw structures for the following heterocyclic amines:

(a) 5-Methoxyindole

(b) 1,3-Dimethylpyrrole

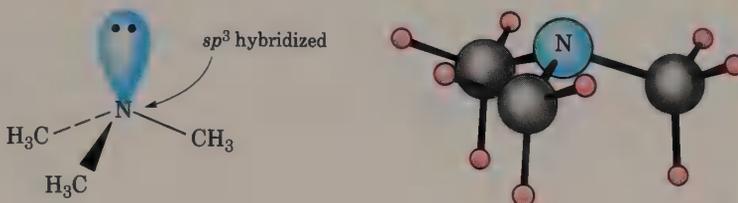
(c) 4-(*N,N*-Dimethylamino)pyridine

(d) 5-Aminopyrimidine

## .....

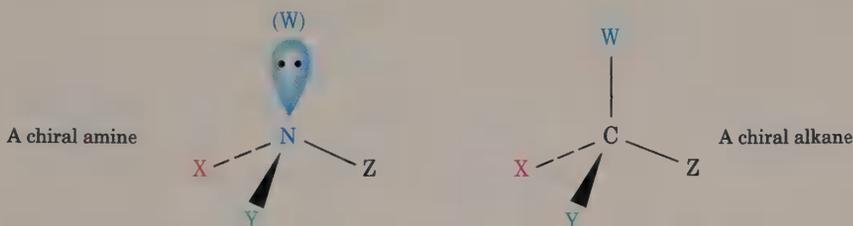
## 25.2 Structure and Bonding in Amines

The bonding in amines is similar to the bonding in ammonia. The nitrogen atom is  $sp^3$  hybridized, with the three substituents occupying three corners of a tetrahedron and the nitrogen's nonbonding lone pair of electrons occupying the fourth corner. As expected, the C–N–C bond angles are very close to the  $109^\circ$  tetrahedral value. For trimethylamine, the C–N–C bond angle is  $108^\circ$  and the C–N bond length is  $1.47 \text{ \AA}$ .



Trimethylamine

One consequence of tetrahedral geometry is that amines with three different substituents on nitrogen are chiral. Such an amine has no plane of symmetry and therefore is not superimposable on its mirror image. If we consider the lone pair of electrons to be the fourth substituent on nitrogen, these chiral amines are analogous to chiral alkanes with four different substituents attached to carbon:

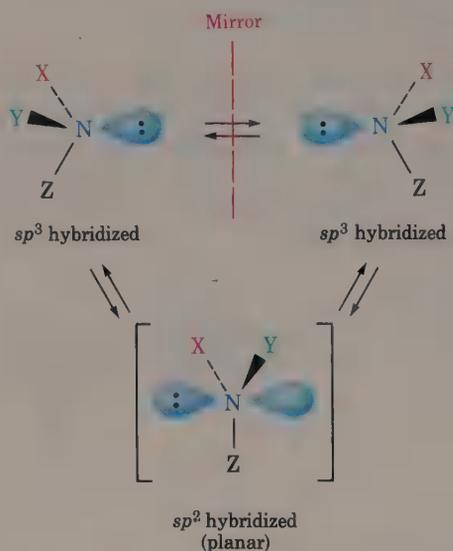


Unlike chiral alkanes, though, most chiral amines can't be resolved into their two enantiomers, because the two enantiomeric forms rapidly interconvert by a **pyramidal inversion**, much as an alkyl halide inverts in an  $S_N2$  reaction. Pyramidal inversion occurs by a momentary rehybridization of the nitrogen atom to planar,  $sp^2$  geometry, followed by rehybridization of the planar intermediate to tetrahedral,  $sp^3$  geometry (Figure 25.1).

Spectroscopic studies have shown that the barrier to nitrogen inversion is about  $6 \text{ kcal/mol}$  ( $25 \text{ kJ/mol}$ ), a figure only twice as large as the barrier to rotation about a carbon–carbon single bond. Pyramidal inversion is therefore rapid at room temperature, and the two optically active forms can't normally be isolated.

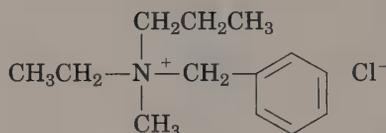
### PROBLEM .....

- 25.5** Although rapid pyramidal inversions prevent chiral trialkylamines from being resolved, chiral tetraalkylammonium salts such as *N*-ethyl-*N*-methyl-*N*-propylbenzylammonium chloride are configurationally stable. Draw both the *R* and the *S*



**Figure 25.1** Pyramidal inversion of amines interconverts the two mirror-image forms.

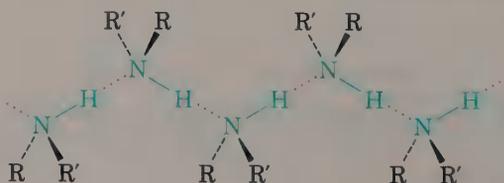
enantiomers of this chiral ammonium salt, and explain why they don't interconvert by pyramidal inversion.



*N*-Ethyl-*N*-methyl-*N*-propylbenzylammonium chloride

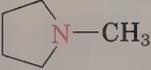
## 25.3 Physical Properties of Amines

Amines are highly polar and therefore have higher boiling points than alkanes of equivalent molecular weight have. Like alcohols, amines with fewer than five carbon atoms are generally water-soluble. Also like alcohols, primary and secondary amines form strong hydrogen bonds and are highly associated in the liquid state. The physical properties of some simple amines are given in Table 25.1.



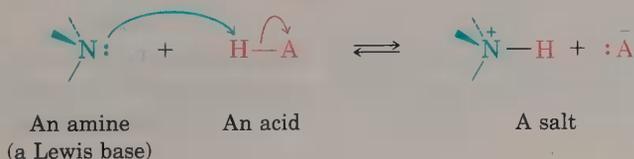
One characteristic of amines that doesn't show up in Table 25.1 is their *odor*. All low-molecular-weight amines have a characteristic and distinctive fish-like aroma, while diamines such as putrescine (1,4-butanediamine) have names that are self-explanatory.

**TABLE 25.1** Physical Properties of Some Simple Amines

Name	Structure	Melting point (°C)	Boiling point (°C)
Ammonia	NH <sub>3</sub>	-77.7	-33.3
<b>Primary amines</b>			
Methylamine	CH <sub>3</sub> NH <sub>2</sub>	-94	-6.3
Ethylamine	CH <sub>3</sub> CH <sub>2</sub> NH <sub>2</sub>	-81	16.6
<i>tert</i> -Butylamine	(CH <sub>3</sub> ) <sub>3</sub> CNH <sub>2</sub>	-67.5	44.4
Aniline (an arylamine)	C <sub>6</sub> H <sub>5</sub> NH <sub>2</sub>	-6.3	184.1
<b>Secondary amines</b>			
Dimethylamine	(CH <sub>3</sub> ) <sub>2</sub> NH	-93	7.4
Diethylamine	(CH <sub>3</sub> CH <sub>2</sub> ) <sub>2</sub> NH	-48	56.3
Diisopropylamine	[(CH <sub>3</sub> ) <sub>2</sub> CH] <sub>2</sub> NH	-61	84
Pyrrolidine		2	89
<b>Tertiary amines</b>			
Trimethylamine	(CH <sub>3</sub> ) <sub>3</sub> N	-117	3
Triethylamine	(CH <sub>3</sub> CH <sub>2</sub> ) <sub>3</sub> N	-114	89.3
<i>N</i> -Methylpyrrolidine		-21	81

## 25.4 Amine Basicity

The chemistry of amines is dominated by the nitrogen lone pair of electrons. Because of the nitrogen lone pair, amines are both basic and nucleophilic. They react with Lewis acids to form acid/base salts, and they react with electrophiles in many of the polar reactions seen in past chapters.



Amines are much more basic than alcohols, ethers, or water. When an amine is dissolved in water, an equilibrium is established in which water acts as a protic acid and donates a proton to the amine. Not all amines are equal in base strength, though; some amines are stronger and some are weaker. Just as we were able to measure the acid strength of a carboxylic acid (Section 20.3), we can also measure the base strength of an amine.

The most convenient way to measure the *basicity* of an amine is to look at the *acidity* ( $pK_a$ ) of the corresponding ammonium salt. Since a more strongly basic amine holds a proton more tightly, its corresponding ammonium ion is less acidic (higher  $pK_a$ ). A more weakly basic amine, however, holds a proton less tightly, and its corresponding ammonium ion is more acidic (lower  $pK_a$ ). This correlation between acidity and basicity is, of course, just another example of the general relationship we saw in Section 2.6. The conjugate base ( $\text{RNH}_2$ ) that results from deprotonation of a weak acid ( $\text{RNH}_3^+$ ) is a strong base, and the conjugate base from deprotonation of a strong acid is a weak base.

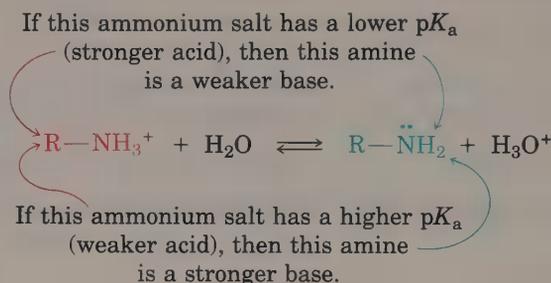
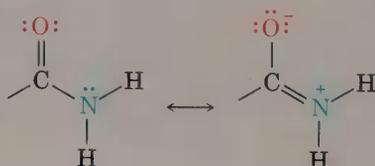


Table 25.2 lists the measured  $pK_a$ 's of some common ammonium salts and indicates that there is little effect of substitution on alkylamine basicity. The salts of most simple alkylamines have  $pK_a$ 's in the narrow range, 10–11, regardless of their exact structure.

In contrast to amines, *amides* ( $\text{RCONH}_2$ ) are completely nonbasic. Amides don't form salts when treated with aqueous acids; their aqueous solutions are neutral; and they are very poor nucleophiles. There are two main reasons for the difference in basicity between amines and amides. First, the ground state of an amide is stabilized by delocalization of the nitrogen lone-pair electrons through orbital overlap with the carbonyl group. In resonance terms, we can draw two contributing forms:

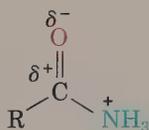


Since this amide resonance stabilization is lost in the protonated product, protonation is disfavored (larger positive  $\Delta G^\circ$ ). Second, a protonated amide is higher in energy than a protonated amine because the electron-with-

TABLE 25.2 Basicity of Some Common Alkylamines

Name	Structure	$pK_a$ of ammonium ion
Ammonia	$:\text{NH}_3$	9.26
<b>Primary alkylamine</b>		
Methylamine	$\text{CH}_3\ddot{\text{N}}\text{H}_2$	10.64
Ethylamine	$\text{CH}_3\text{CH}_2\ddot{\text{N}}\text{H}_2$	10.75
<b>Secondary alkylamine</b>		
Dimethylamine	$(\text{CH}_3)_2\ddot{\text{N}}\text{H}$	10.73
Diethylamine	$(\text{CH}_3\text{CH}_2)_2\ddot{\text{N}}\text{H}$	10.94
Pyrrolidine		11.27
<b>Tertiary alkylamine</b>		
Trimethylamine	$(\text{CH}_3)_3\text{N}:$	9.79
Triethylamine	$(\text{CH}_3\text{CH}_2)_3\text{N}:$	10.75

drawing carbonyl group inductively destabilizes the neighboring positive charge.

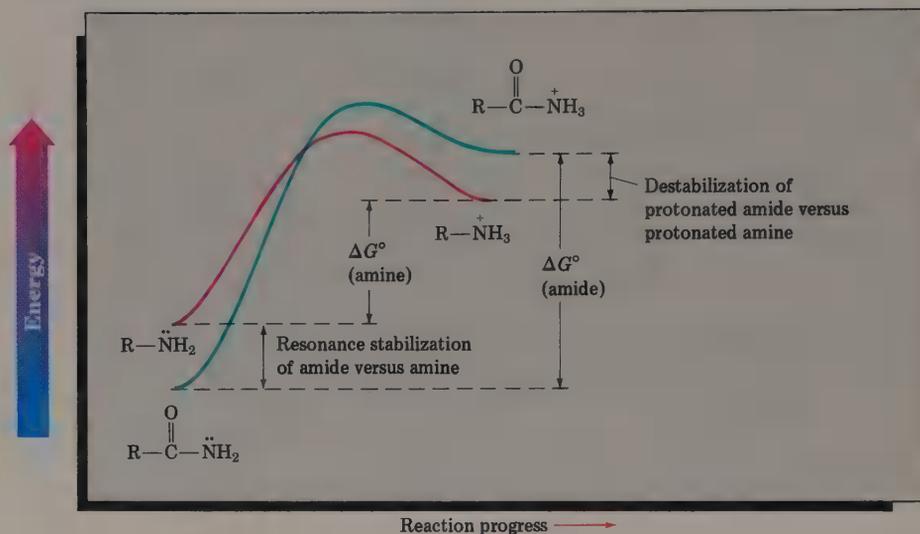


A protonated amide has no resonance stabilization and has inductive destabilization of the positive charge.

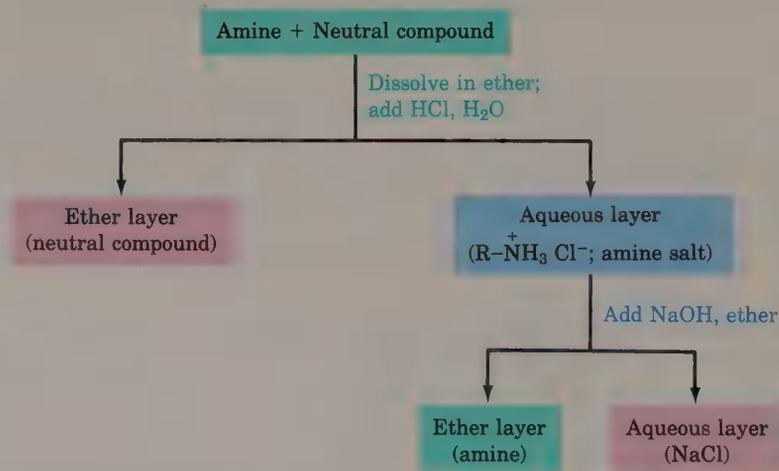
Both factors—increased stability of an amide versus an amine and decreased stability of a protonated amide versus a protonated amine—lead to a large difference in  $\Delta G^\circ$  and a resultant large difference in basicity for amines and amides. Figure 25.2 shows these relationships on a reaction energy diagram.

It's often possible to take advantage of their basicity to purify amines. For example, if a mixture of a basic amine and a neutral compound such as a ketone, an alcohol, or an ether is dissolved in an organic solvent and extracted with aqueous acid, the basic amine dissolves in the water layer as its protonated salt, while the neutral compound remains in the organic solvent layer. Separation, basification, and extraction of the aqueous layer with organic solvent then provides the pure amine (Figure 25.3).

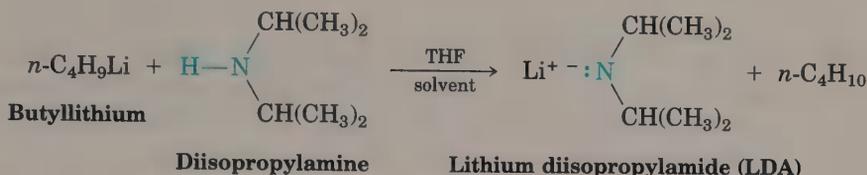
In addition to their behavior as bases, primary and secondary amines can also act as extremely weak *acids*, since their N–H protons can be removed by a sufficiently strong base. We've already seen, for example, how diisopropylamine ( $pK_a \approx 40$ ) reacts with butyllithium to yield lithium diisopropylamide (LDA) and butane (Section 22.5).



**Figure 25.2** A reaction energy diagram comparing the protonation of amines and amides. The  $\Delta G^\circ$  for amide protonation is larger because of increased reactant stability and decreased product stability.



**Figure 25.3** Separation and purification of an amine component from a mixture.



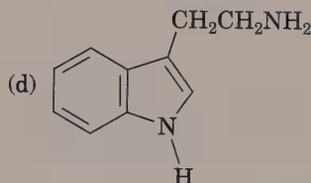
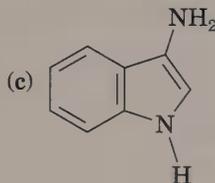
Dialkylamide anions like LDA are extremely powerful bases that are much used in organic chemistry, particularly for the generation of enolate ions from carbonyl compounds (Section 22.9).

## PROBLEM

- 25.6 Which compound in each of the following pairs is more basic?  
 (a)  $\text{CH}_3\text{CH}_2\text{NH}_2$  or  $\text{CH}_3\text{CH}_2\text{CONH}_2$       (b)  $\text{NaOH}$  or  $\text{CH}_3\text{NH}_2$   
 (c)  $\text{CH}_3\text{NHCH}_3$  or  $\text{CH}_3\text{OCH}_3$

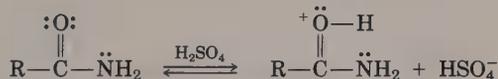
## PROBLEM

- 25.7 In each of the following molecules, predict which nitrogen atom is more basic.  
 (a) Imidazole      (b) 2-Aminoquinoline



## PROBLEM

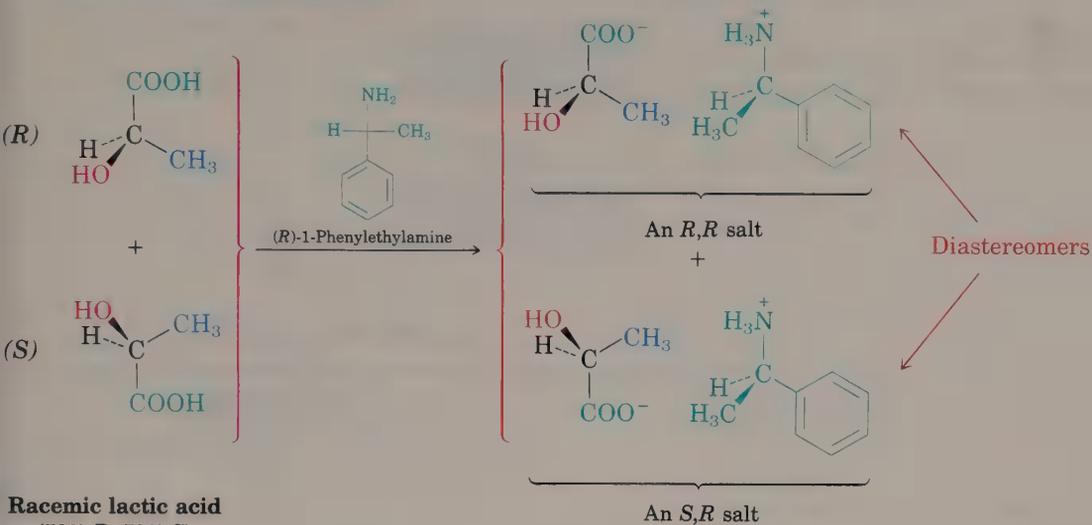
- 25.8 Protonation of an amide actually occurs on oxygen rather than on nitrogen. Suggest a reason for this behavior.



## 25.5 Resolution of Enantiomers via Amine Salts

We saw in the previous section that it's possible to take advantage of the basic properties of an amine to carry out its purification. We can also take advantage of amine basicity in another important way: to carry out the **resolution** of a racemic carboxylic acid into its two pure enantiomers. [Recall from Section 9.10 that a racemic mixture is a 50:50 mixture of (+) and (−) enantiomers.]





**Figure 25.5** Reaction of racemic lactic acid with pure (*R*)-1-phenylethylamine yields a mixture of diastereomeric ammonium salts.

In the same way, (+)- and (–)-lactic acids react with (*R*)-1-phenylethylamine to give two different products. (*R*)-Lactic acid reacts with (*R*)-1-phenylethylamine to give the *R,R* salt, whereas (*S*)-lactic acid reacts with the same *R* amine to give the *S,R* salt. *These two salts are diastereomers.* They are different compounds and have different chemical and physical properties. It may therefore prove possible to separate them physically by fractional crystallization or by some other laboratory technique. Once separated, acidification of the two diastereomeric salts with mineral acid then allows us to isolate the two pure enantiomers of lactic acid and to recover the pure chiral amine for further use. A flow diagram for the overall process is shown in Figure 25.6.

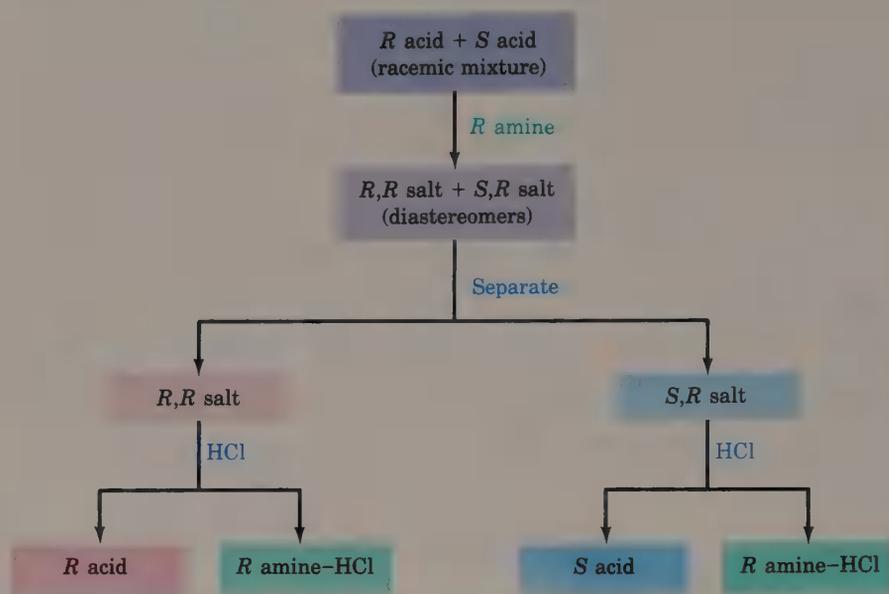
The reaction between two enantiomers of a chiral acid and one enantiomer of a chiral base to yield diastereomeric salts is an example of a general rule: *Any reaction between a racemic mixture and another chiral partner always yields a mixture of diastereomeric products.*

**PROBLEM** .....

- 25.9** Suppose that racemic lactic acid reacts with methanol to yield the ester methyl lactate. What stereochemistry would you expect the product(s) to have? What is the relationship of one product to another?

**PROBLEM** .....

- 25.10** Suppose that a chiral acid such as (*S*)-lactic acid reacts with a chiral alcohol such as (*R*)-2-butanol to form an ester. What stereochemistry would you expect the product(s) to have? Draw the starting materials and product.



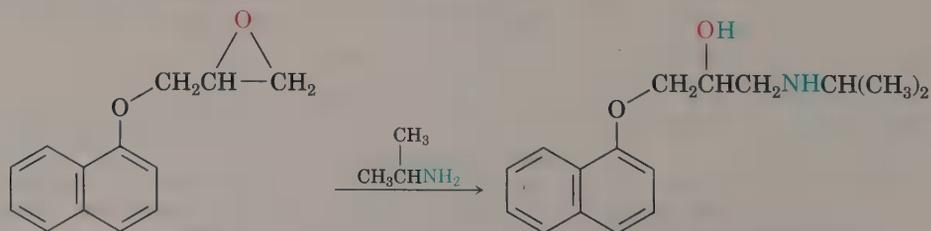
**Figure 25.6** A flow diagram for the resolution of a racemic carboxylic acid into its two pure enantiomers.

**PROBLEM** .....

- 25.11** Suppose that racemic lactic acid reacts with (*S*)-2-butanol to form an ester. What stereochemistry does the product(s) have? What is the relationship of one product to another? How might you use this reaction to resolve ( $\pm$ )-lactic acid?
- .....

## 25.6 Industrial Sources and Uses of Alkylamines

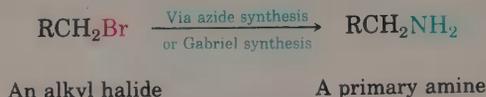
Simple alkylamines such as methylamine, ethylamine, and dimethylamine all find a variety of minor applications in the chemical industry as starting materials for the preparation of insecticides and pharmaceuticals. For example, propranolol, a heart stimulant used in the control of cardiac arrhythmias, is prepared by  $S_N2$  reaction of an epoxide with isopropylamine.



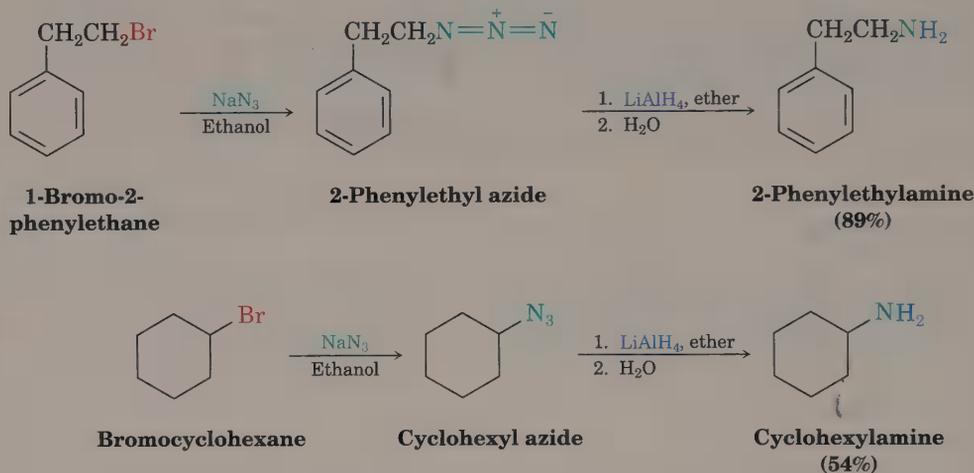
**Propranolol (a heart stimulant)**



used: the **azide synthesis** and the **Gabriel synthesis**. Both methods are excellent ways to prepare primary amines from alkyl halides:

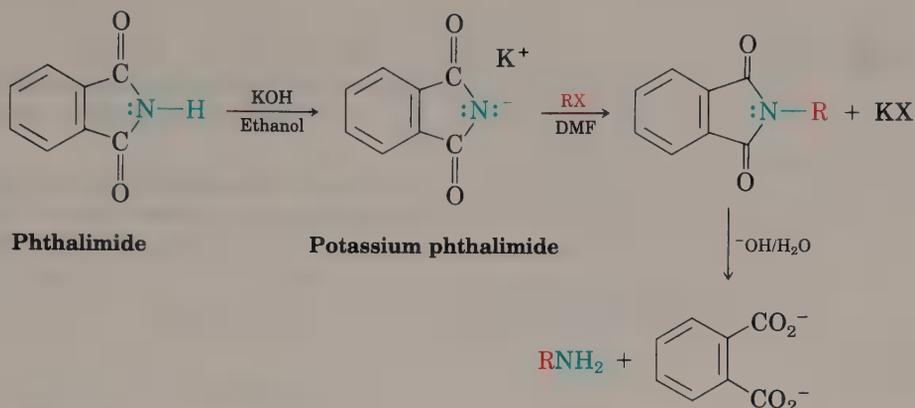


Azide ion,  $\text{N}_3^-$ , is a nonbasic, highly reactive nucleophile that displaces halide ion from primary and even secondary alkyl halides to give alkyl azides,  $\text{RN}_3$ , in high yield. Since alkyl azides are not themselves nucleophilic, overalkylation can't occur. Reduction of alkyl azides, either by catalytic hydrogenation over palladium or by reaction with  $\text{LiAlH}_4$ , leads to the desired primary amine. Yields are usually excellent, but the value of the process is tempered by the fact that low-molecular-weight alkyl azides are explosive and must be handled carefully.

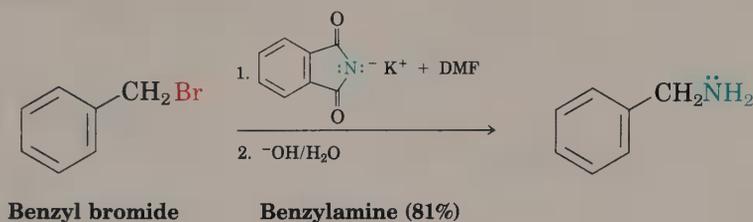


The Gabriel<sup>1</sup> amine synthesis, introduced in 1887, uses an imide alkylation to provide an alternative means of preparing primary amines from alkyl halides. **Imides** ( $-\text{CONHCO}-$ ) are structurally similar to ethyl acetoacetate in that the proton on nitrogen is flanked by two acidifying carbonyl groups. Thus, imides are readily deprotonated by such bases as  $\text{KOH}$ , and their resultant anions are readily alkylated in a reaction similar to the acetoacetic ester synthesis (Section 22.9). Basic hydrolysis of the *N*-alkyl phthalimide then yields a primary amine product. The reaction is best carried out in the highly polar solvent, dimethylformamide, DMF [ $\text{HCON}(\text{CH}_3)_2$ ]. Note that the imide hydrolysis step to yield amine plus phthalate ion is closely analogous to the hydrolysis of an amide (Section 21.8).

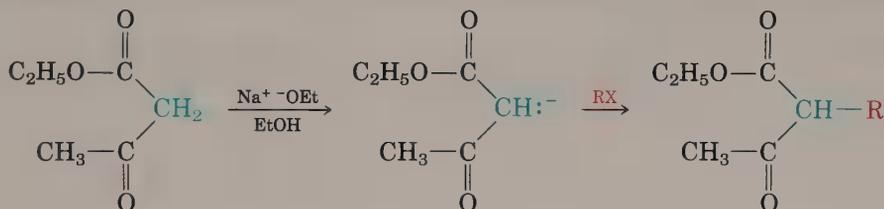
<sup>1</sup>Siegmond Gabriel (1851–1924); b. Berlin; Ph.D. University of Berlin (1874); assistant to A. W. von Hofmann; professor, University of Berlin.



For example,



Recall the acetoacetic ester synthesis:

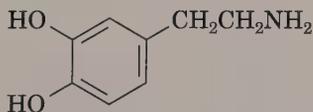


PROBLEM .....

- 25.12 Show the mechanism of the last step in the Gabriel amine synthesis, the base-promoted hydrolysis of an imide to yield an amine plus phthalate ion.

PROBLEM .....

- 25.13 Show two methods for the synthesis of dopamine, a neurotransmitter involved in regulation of the central nervous system. Use any alkyl halide needed.



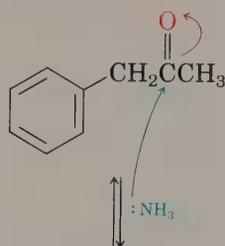
**Dopamine**

.....

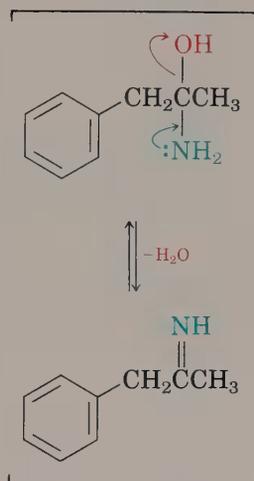


Reductive amination takes place by the pathway shown in Figure 25.7. As indicated, an imine intermediate is first formed by a nucleophilic addition reaction (Section 19.12), and the imine is then reduced.

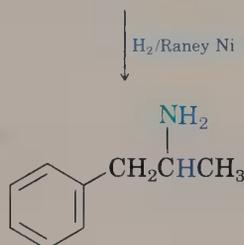
Ammonia attacks the carbonyl group in a nucleophilic addition reaction to yield an intermediate carbinolamine.



The intermediate loses water to give an imine.

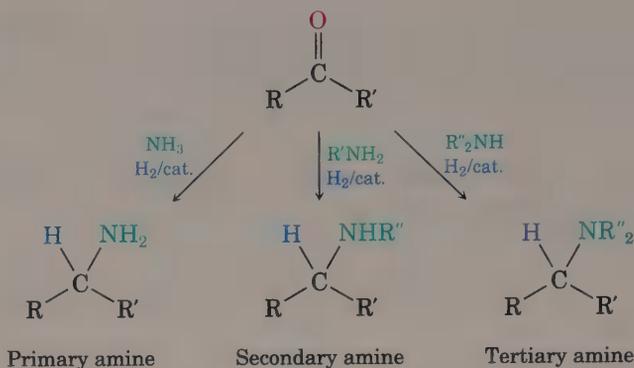


The imine is reduced catalytically over Raney nickel to yield the amine product.

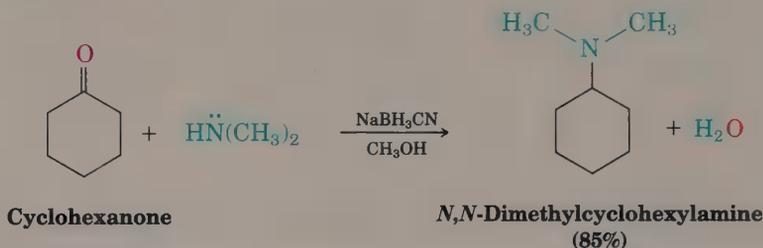


**Figure 25.7** Mechanism of reductive amination of a ketone to yield an amine. (Complete details of the imine-forming step are shown in Figure 19.8.)

Ammonia, primary amines, and secondary amines can all be used in the reductive amination reaction to yield primary, secondary, and tertiary amines, respectively.



Many different reducing agents are effective, but the most common choice on a laboratory scale (as opposed to an industrial scale) is sodium cyanoborohydride,  $\text{NaBH}_3\text{CN}$ .



PROBLEM .....

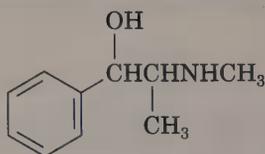
- 25.15 Show how these amines might be prepared by reductive amination of a ketone or an aldehyde. Show all precursors if more than one is possible.
- (a)  $\text{CH}_3\text{CH}_2\text{NHCH}(\text{CH}_3)_2$                       (b) *N*-Ethylaniline  
 (c) *N*-Methylcyclopentylamine

PROBLEM .....

- 25.16 Show the mechanism of reductive amination of cyclohexanone and dimethylamine with  $\text{NaBH}_3\text{CN}$ . What intermediates are involved?

PROBLEM .....

- 25.17 Ephedrine is an amino alcohol that is widely used for the treatment of bronchial asthma. Show how a reductive amination step might be used to synthesize ephedrine.

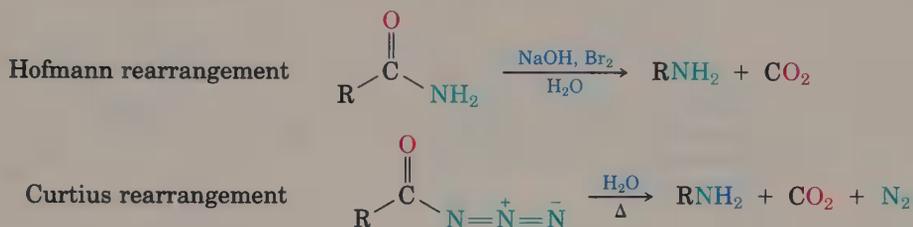


Ephedrine

.....

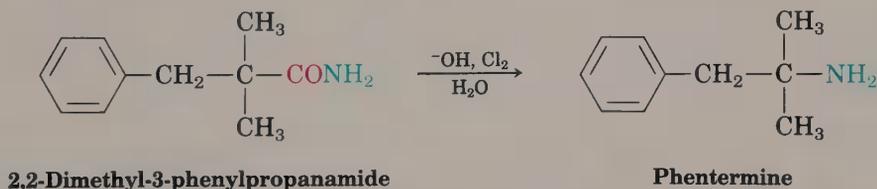
## HOFMANN AND CURTIUS REARRANGEMENTS

Carboxylic acid derivatives can be converted into primary amines with loss of one carbon atom ( $\text{RCOY} \rightarrow \text{RNH}_2$ ) by both the **Hofmann<sup>2</sup> rearrangement** and the **Curtius<sup>3</sup> rearrangement**. Although the Hofmann rearrangement involves a primary amide and the Curtius rearrangement involves an acyl azide, both proceed through similar mechanisms:



Hofmann rearrangement occurs when a primary amide,  $\text{RCONH}_2$ , is treated with halogen and base (Figure 25.8). Although the overall mechanism is lengthy, most of the individual steps have been encountered before. Thus, the bromination of an amide in steps 1 and 2 is analogous to the base-promoted bromination of a ketone enolate ion (Section 22.7), and the rearrangement of the bromoamide anion in step 4 is analogous to a carbocation rearrangement (Section 6.13). The primary difference between the migration step in a Hofmann rearrangement and that in a carbocation rearrangement is that the  $\text{R}^-$  group begins its migration to the neighboring atom *at the same time* that the bromide ion is leaving, rather than after it has left. Nucleophilic addition of water to the isocyanate carbonyl group in step 5 is a typical carbonyl-group process (Section 19.8), as is the final decarboxylation step (Section 22.9).

Despite its mechanistic complexity, the Hofmann rearrangement often gives high yields of both aryl- and alkylamines. For example, the appetite-suppressing drug phentermine is prepared commercially by Hofmann rearrangement of a primary amide:

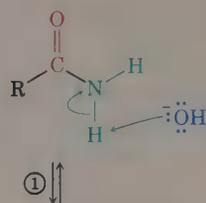


The Curtius rearrangement, like the Hofmann rearrangement, involves migration of an  $\text{R}^-$  group from the carbonyl carbon atom to the neighboring nitrogen with simultaneous loss of a leaving group. The reaction takes place

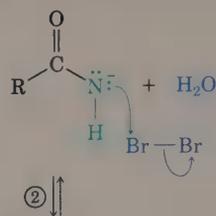
<sup>2</sup>August Wilhelm von Hofmann (1818–1892); b. Giessen, Germany; professor, Bonn, the Royal College of Chemistry, London (1845–1864), Berlin (1865–1892).

<sup>3</sup>Theodor Curtius (1857–1928); b. Duisberg; Ph.D. Leipzig; professor, universities of Kiel, Bonn, and Heidelberg (1898–1926).

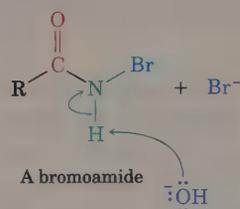
Base abstracts an acidic N-H proton, yielding an anion.



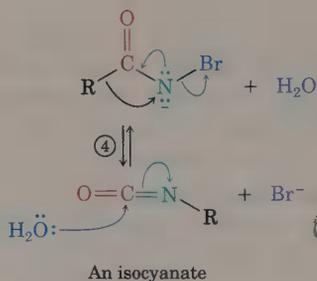
The anion reacts with bromine in an alpha-substitution reaction to give an *N*-bromoamide.



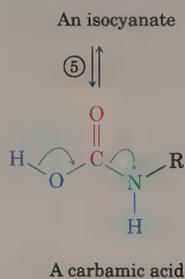
Base abstraction of the remaining amide proton gives a bromoamide anion.



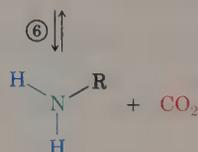
The bromoamide anion rearranges as the R group attached to the carbonyl carbon migrates to nitrogen at the same time the bromide ion leaves, giving an isocyanate.



The isocyanate adds water in a nucleophilic addition step to yield a carbamic acid.

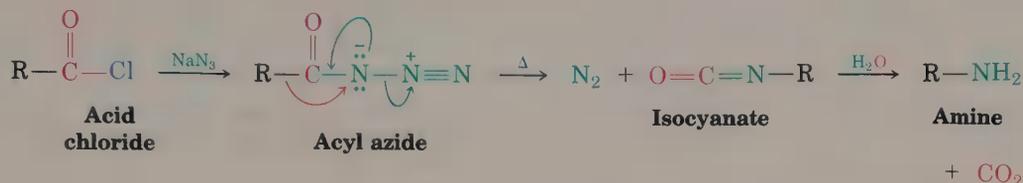


The carbamic acid spontaneously loses CO<sub>2</sub>, yielding the amine product.

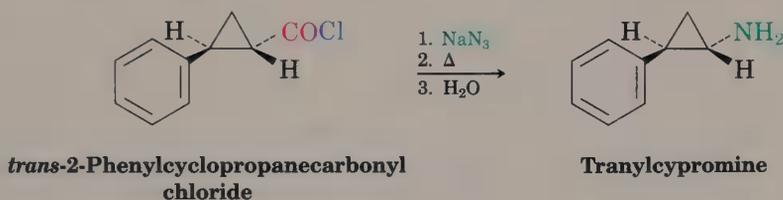


**Figure 25.8** Mechanism of the Hofmann rearrangement of an amide to an amine. Each step is analogous to a reaction studied previously.

on heating an acyl azide that is itself prepared by nucleophilic acyl substitution of an acid chloride.



Like the Hofmann rearrangement, the Curtius rearrangement is often used commercially. For example, the antidepressant drug tranlycypromine is made by Curtius rearrangement of 2-phenylcyclopropanecarbonyl chloride:



PROBLEM .....

- 25.18 Show the mechanism of the Curtius rearrangement of an acyl azide to an isocyanate, indicating the origin and fate of all bonding electrons. Show also the mechanism of the addition of water to an isocyanate to yield a carbamic acid.

PROBLEM .....

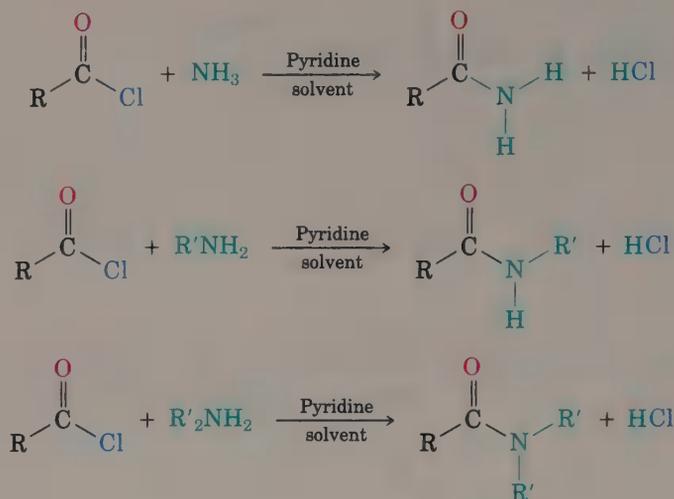
- 25.19 What starting materials would you use to prepare these amines by Hofmann and Curtius rearrangements?



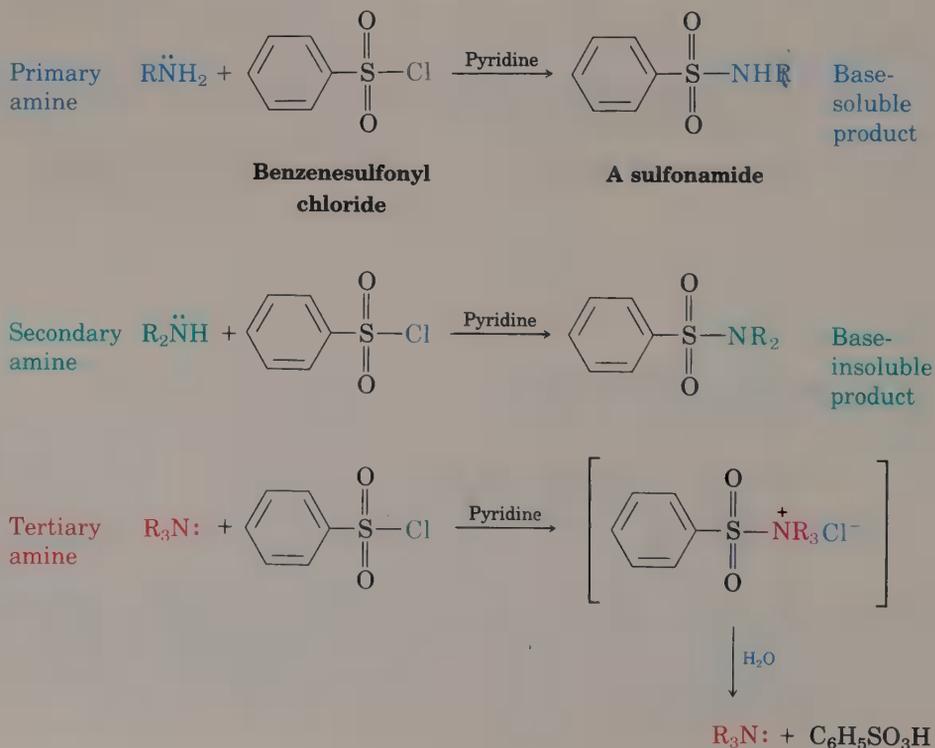
.....

## 25.8 Reactions of Amines

We've already studied the two most important reactions of alkylamines: alkylation and acylation. As we saw earlier in this chapter, primary, secondary, and tertiary amines can all be alkylated by reaction with primary alkyl halides. Alkylations of primary and secondary amines are difficult to control and often give mixtures of products, but tertiary amines are cleanly alkylated to give quaternary ammonium salts. Primary and secondary (but not tertiary) amines can also be acylated by reaction with acid chlorides or acid anhydrides (Sections 21.5 and 21.6) to yield amides.

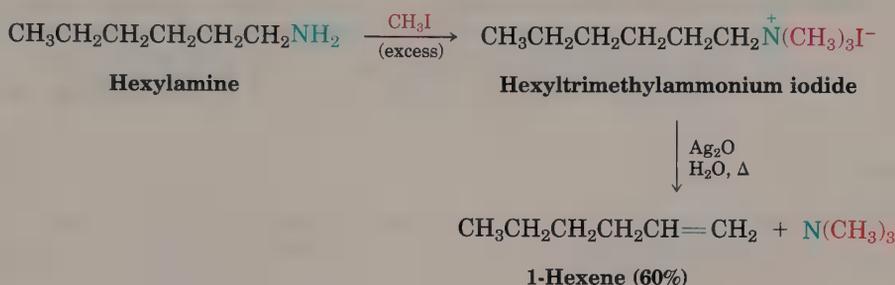


If a **sulfonyl chloride**,  $\text{R}'\text{SO}_2\text{Cl}$ , is used as the acylating agent, a **sulfonamide** ( $\text{R}'\text{SO}_2\text{NR}_2$ ) is produced. This reaction forms the basis of the **Hinsberg test**, a classic laboratory method for distinguishing among primary, secondary, and tertiary amines. A primary amine yields a sulfonamide product that has one remaining acidic N-H proton and can therefore be identified by its solubility in aqueous NaOH. A secondary amine yields a sulfonamide that has no acidic protons and can be identified by its lack of base solubility. A tertiary amine yields a product that reacts instantly with water to give back amine starting material. Thus, no apparent reaction occurs with a tertiary amine.

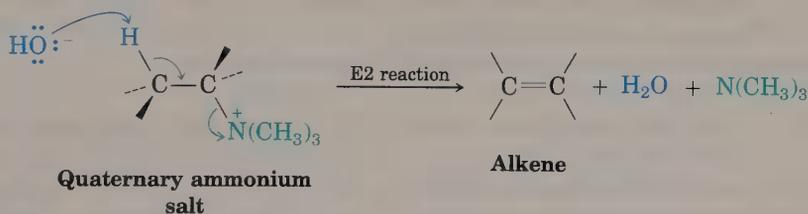


## HOFMANN ELIMINATION

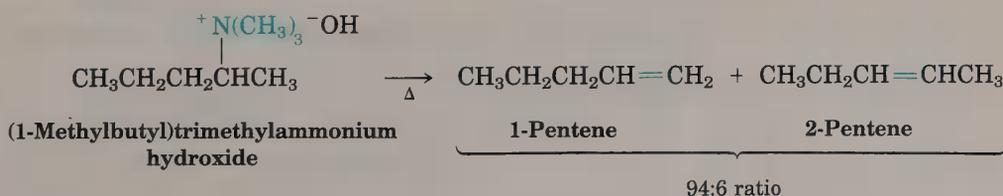
Just as alcohols can be dehydrated to yield alkenes, amines can also be converted into alkenes under suitable conditions. In the **Hofmann elimination reaction**, an amine is first methylated with excess iodomethane to produce a quaternary ammonium salt, which then undergoes an elimination reaction to give an alkene on heating with silver oxide. For example, 1-hexene is formed from hexylamine in 60% yield:



Silver oxide functions by exchanging hydroxide ion for iodide ion in the quaternary salt, thus providing the base necessary to effect elimination. The actual elimination step is an E2 reaction (Section 11.11) in which hydroxide ion removes a proton and the positively charged nitrogen atom acts as the leaving group.



An interesting feature of the Hofmann elimination is that it gives products different from those of most other E2 reactions. Whereas the *more* highly substituted alkene product generally predominates in the E2 reaction of an alkyl halide (Zaitsev's rule; Section 11.10), the *less* highly substituted alkene predominates in the Hofmann elimination of a quaternary ammonium salt. The reasons for this selectivity aren't fully understood, but are probably steric in origin. Because of the large size of the trialkylamine leaving group, hydroxide ion must abstract a hydrogen from the most sterically accessible, least hindered position. For example, (1-methylbutyl)trimethylammonium hydroxide yields 1-pentene and 2-pentene in a 94:6 ratio:



The Hofmann elimination reaction is important primarily because of its historical use as a degradative tool in the structure determination of many complex naturally occurring amines. The reaction is not often used today, since the product alkenes can usually be made more easily in other ways.

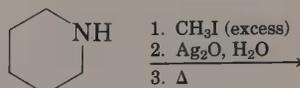
## PROBLEM .....

**25.20** What products would you expect to obtain from Hofmann elimination of these amines? If more than one product is formed, indicate which is major.

- (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{NH}_2)\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$       (b) Cyclohexylamine  
 (c)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{NH}_2)\text{CH}_2\text{CH}_2\text{CH}_3$       (d) *N*-Ethylcyclohexylamine

## PROBLEM .....

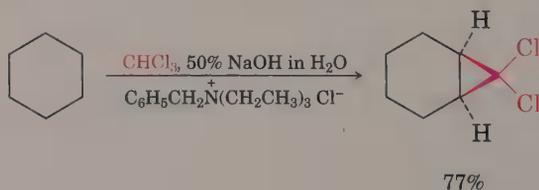
**25.21** What product would you expect from Hofmann elimination of a cyclic amine such as piperidine? Formulate all the steps involved.



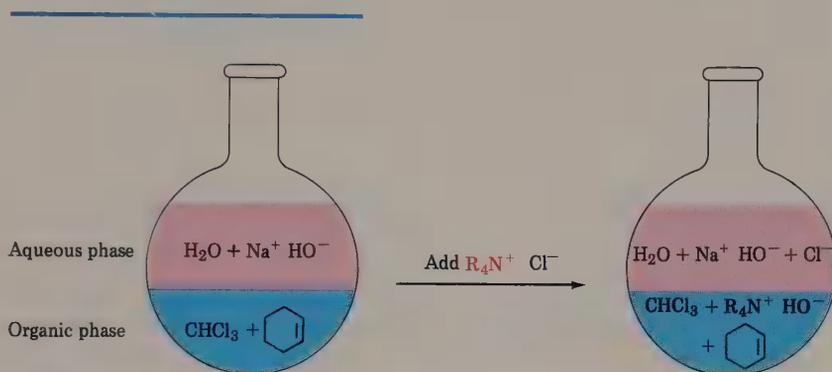
## 25.9 Tetraalkylammonium Salts as Phase-Transfer Agents

Tetraalkylammonium salts,  $\text{R}_4\text{N}^+ \text{X}^-$ , are easily prepared by  $\text{S}_{\text{N}}2$  reaction between a tertiary amine and an alkyl halide and have come to be widely used in the past decade as catalysts for many different kinds of organic reactions. For example, we saw in Section 7.11 that chloroform reacts with a strong base such as NaOH to generate dichlorocarbene, which can then add to a carbon-carbon double bond to yield a dichlorocyclopropane.

Imagine an experiment where cyclohexene is dissolved in chloroform and the organic solution is stirred with 50% aqueous sodium hydroxide. Since the organic layer and the water layer are immiscible, the strong base in the aqueous phase is unable to come into contact with chloroform in the organic phase, and there is no reaction. If, however, a small amount of a tetraalkylammonium salt such as benzyltriethylammonium chloride is added to the two-phase mixture, an immediate reaction occurs, leading to formation of the dichlorocyclopropane product in 77% yield:

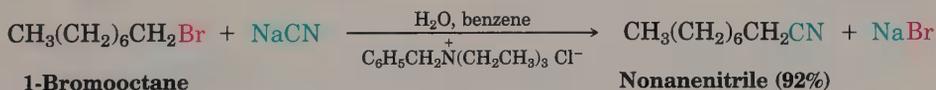


How does the tetraalkylammonium salt exert its catalytic effect? Benzyltriethylammonium ion, even though positively charged, is nevertheless soluble in organic solvents because of the four hydrocarbon substituents on nitrogen. When the positively charged tetraalkylammonium ion goes into the organic layer, a negatively charged counter-ion must also move into the organic layer to preserve charge neutrality. Hydroxide ion, present in far greater amount than chloride ion, is thus transferred into the organic phase where reaction with chloroform immediately occurs (Figure 25.9).



**Figure 25.9** Phase-transfer catalysis. Addition of a small amount of a tetraalkylammonium salt to a two-phase mixture allows an inorganic anion to be transferred from the aqueous phase into the organic phase, where a reaction can occur.

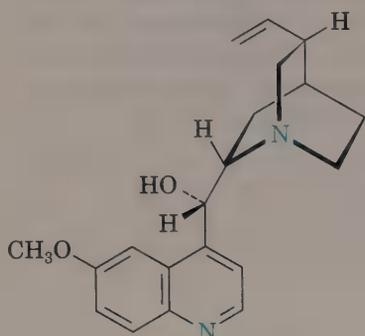
Transfer of an inorganic ion from one phase to another is called **phase transfer**, and the tetraalkylammonium salt is referred to as a **phase-transfer catalyst**. Many different kinds of organic reactions, including oxidations, reductions, carbonyl-group alkylations, and  $S_N2$  reactions, are subject to phase-transfer catalysis, often with considerable improvements in yield.  $S_N2$  reactions are particularly good candidates for phase-transfer catalysis, since inorganic nucleophiles can be transferred from an aqueous (protic) phase to an organic (aprotic) phase, where they are far more reactive. For example:



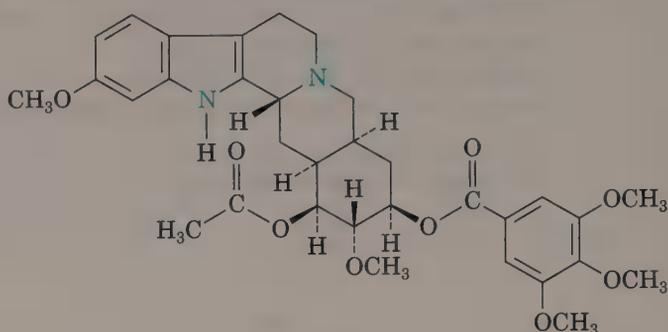
## 25.10 Naturally Occurring Amines: Alkaloids

A large variety of amines are widely distributed among plants and animals. Trimethylamine, for instance, occurs in animal tissues and is partially responsible for the distinctive odor of many fish; quinine is an important

antimalarial drug isolated from the bark of the South American *Cinchona* tree; and reserpine is a useful antihypertensive (blood-pressure-lowering) agent isolated from the Indian shrub *Rauwolfia serpentina*.

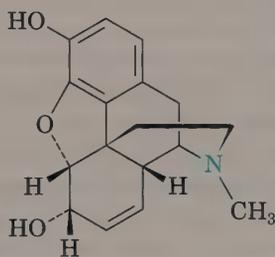


Quinine (antimalarial)

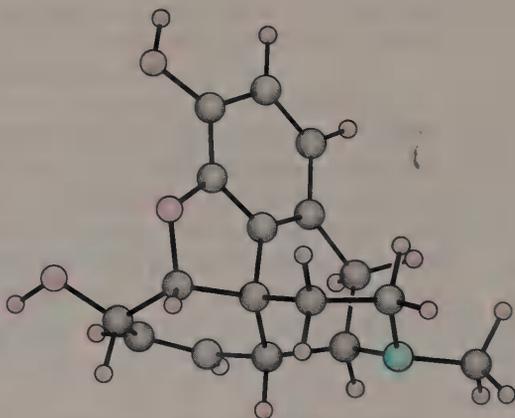


Reserpine (antihypertensive)

Naturally occurring amines derived from plant sources were once known as “vegetable alkali,” since their aqueous solutions are slightly basic, but they are now referred to as **alkaloids**. The study of alkaloids provided much of the impetus for the growth of organic chemistry in the nineteenth century and remains today a fascinating area of research. Rather than attempt to classify the many kinds of alkaloids, let's look briefly at one particular group, the morphine alkaloids.

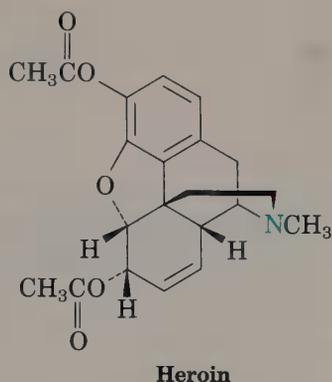
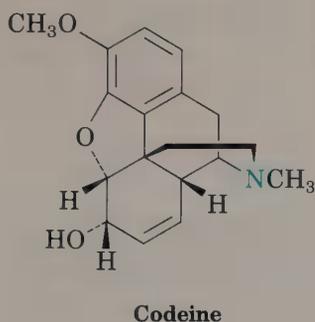


Morphine (an analgesic)



The medical uses of morphine alkaloids have been known at least since the seventeenth century, when crude extracts of the opium poppy, *Papaver somniferum*, were used for the relief of pain. Morphine was the first pure alkaloid to be isolated from the poppy, but its close relative, codeine, also occurs naturally. Codeine, which is simply the methyl ether of morphine, is used in prescription cough medicines. Heroin, another close relative of morphine, does not occur naturally but is synthesized by diacetylation of morphine.

Chemical investigations into the structure of morphine occupied some of the finest chemical minds of the nineteenth and early twentieth centuries, until the puzzle was finally solved by Robert Robinson in 1925. The key reaction used to establish structure was the Hofmann elimination.

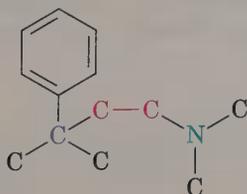


Morphine and its relatives constitute a class of extremely useful pharmaceutical agents, yet they also pose an enormous social problem because of their addictive properties. Much effort has therefore gone into understanding how morphine works and into developing modified morphine analogs that retain the analgesic activity but don't cause physical dependence. Our present understanding is that morphine appears to bind to opiate receptor sites in the brain. It doesn't interfere with or lessen the transmission of a pain signal to the brain but rather changes the brain's reception of the signal.

Hundreds of morphine-like molecules have been synthesized and tested for their analgesic properties. For example, replacement of the morphine *N*-methyl group by an *N*-allyl substituent yields *nalorphine*, an analgesic agent that acts as a narcotic *antagonist* to reverse many of the undesirable side effects of morphine.

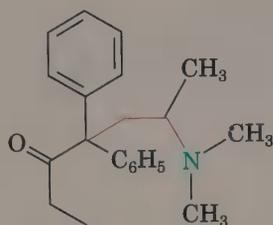


Studies have shown that not all of the complex tetracyclic framework of morphine is necessary for biological activity. According to the "morphine rule," biological activity requires four things: (1) an aromatic ring attached to (2) a quaternary carbon atom and (3) a tertiary amine situated (4) two carbon atoms farther away. Thus, the bicyclic amine meperidine (Demerol) is widely used as a painkiller, and methadone has been used in the treatment of heroin addiction.

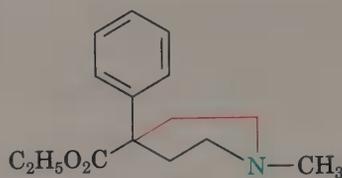


**The morphine rule:**

an aromatic ring, attached to a quaternary carbon, attached to two more carbons, attached to a tertiary amine



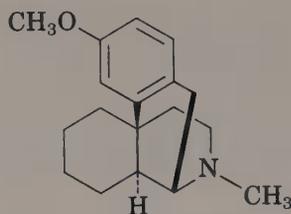
Methadone



Meperidine

## PROBLEM .....

- 25.22 Show how the morphine rule fits the structure of dextromethorphan, a common constituent of cough remedies.



Dextromethorphan

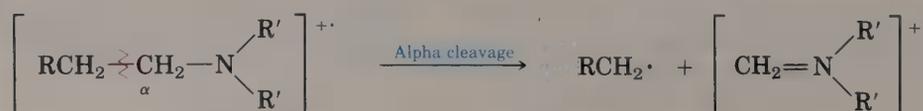
## 25.11 Spectroscopy of Amines

### MASS SPECTROSCOPY

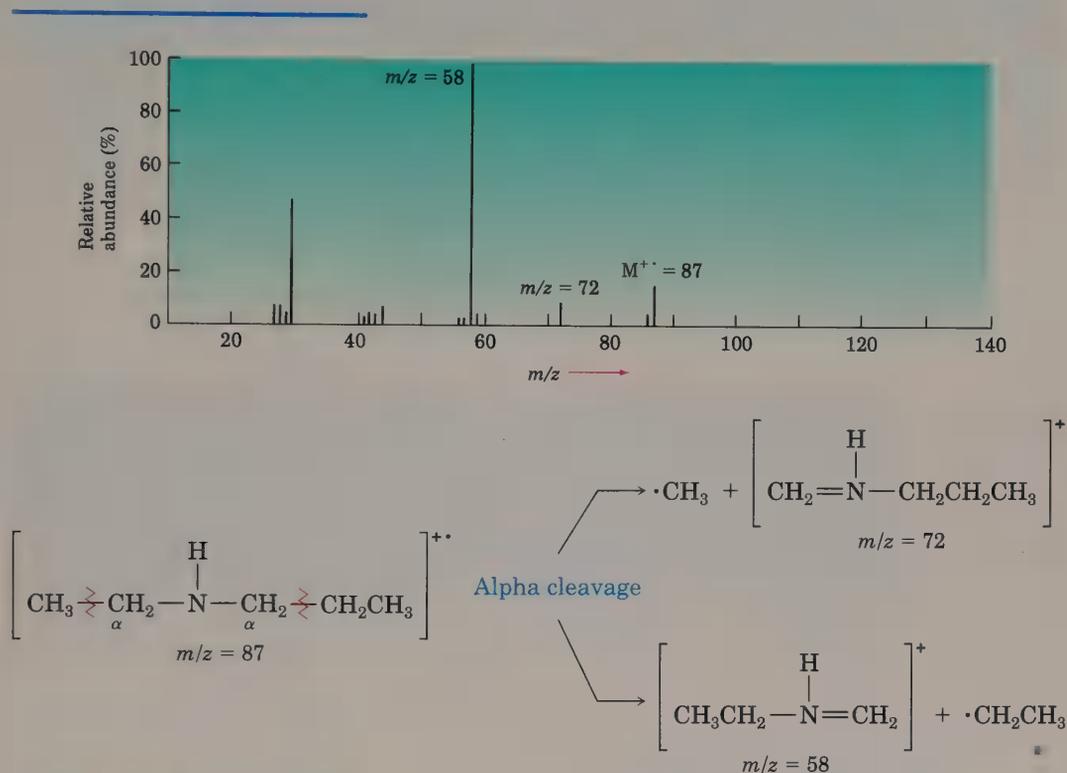
The **nitrogen rule** of mass spectroscopy says that compounds with an odd number of nitrogen atoms have odd-numbered molecular weights. Thus, the presence of nitrogen in a molecule is detected simply by observing its mass spectrum. An odd-numbered molecular ion usually means that the unknown compound has one or three nitrogen atoms, and an even-numbered molecular ion usually means that the compound has either zero or two nitrogen atoms.

The logic behind the rule derives from the fact that nitrogen is trivalent, thus requiring an odd number of hydrogen atoms in the molecule. For example, methylamine has the formula  $\text{CH}_5\text{N}$  and a molecular weight of 31; morphine has the formula  $\text{C}_{17}\text{H}_{19}\text{NO}_3$  and a molecular weight of 285.

Aliphatic amines undergo a characteristic alpha cleavage in the mass spectrometer, similar to the cleavage observed for aliphatic alcohols (Section 17.11). A carbon-carbon bond nearest the nitrogen atom is broken, yielding an alkyl radical and a nitrogen-containing cation:



For example, the mass spectrum of *N*-ethylpropylamine shown in Figure 25.10 has peaks at  $m/z = 58$  and  $m/z = 72$ , corresponding to the two possible modes of alpha cleavage.

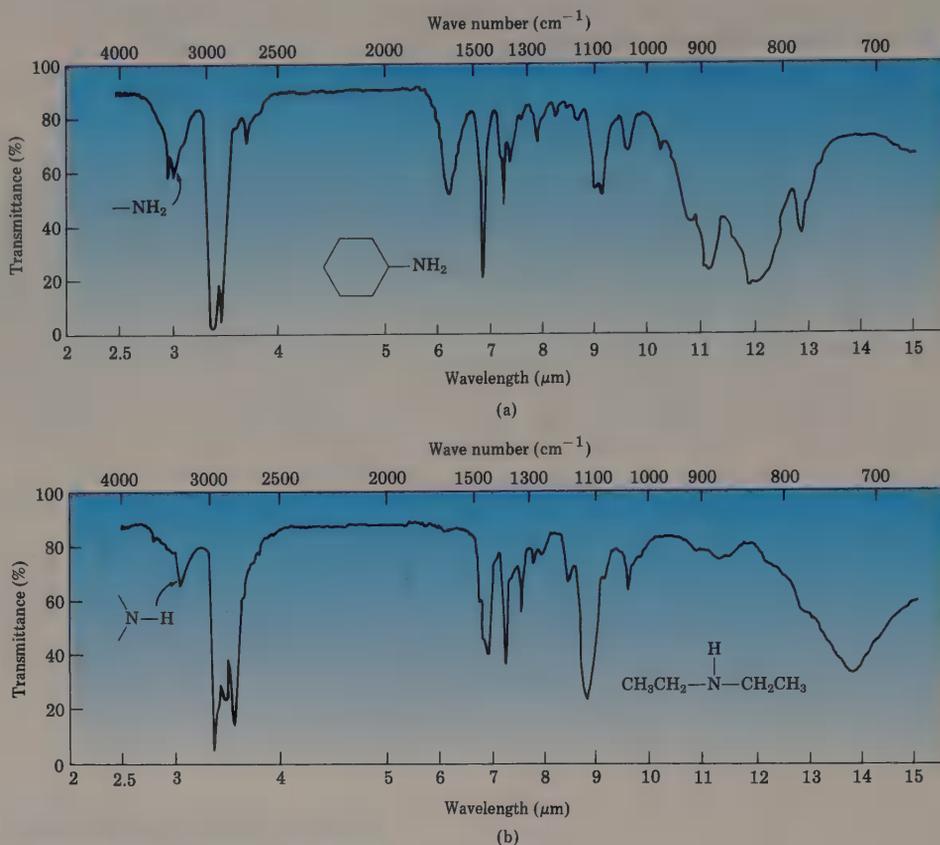


**Figure 25.10** Mass spectrum of *N*-ethylpropylamine. The two possible modes of alpha cleavage occur, leading to the observed fragment ions at  $m/z = 58$  and  $m/z = 72$ .

## INFRARED SPECTROSCOPY

Primary and secondary amines can be identified by characteristic N–H bond stretching absorptions in the  $3300\text{--}3500\text{ cm}^{-1}$  range of the infrared spectrum. Alcohols also absorb in this range (Section 17.11), but amine absorption bands are generally both sharper and less intense than hydroxyl bands. Primary amines show a pair of bands at about  $3400$  and  $3500\text{ cm}^{-1}$ , and secondary amines show a single band at  $3350\text{ cm}^{-1}$ . Tertiary amines show no absorption in this region, since they have no N–H protons. Representative infrared spectra of both primary and secondary amines are shown in Figure 25.11.

In addition to looking for characteristic N–H bands, there's also a simple trick that can be used to tell whether or not a given unknown is an amine. Addition of a small amount of mineral acid produces a broad and strong ammonium band in the  $2200\text{--}3000\text{ cm}^{-1}$  range if the sample contains an

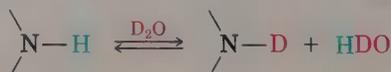


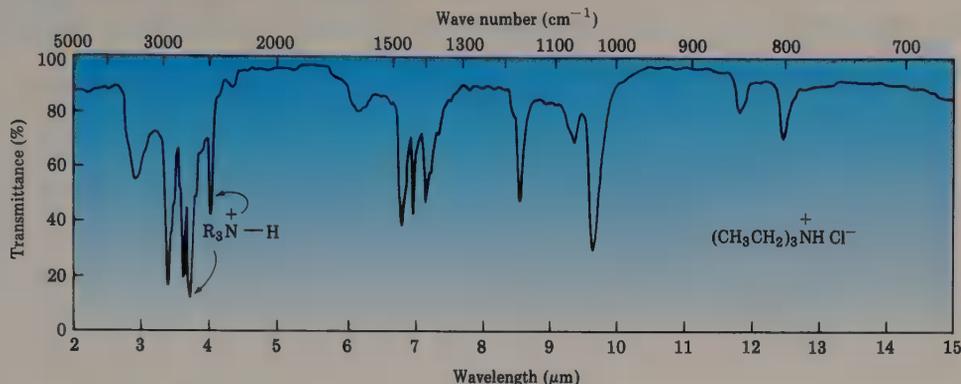
**Figure 25.11** Infrared spectra of (a) cyclohexylamine and (b) diethylamine.

amino group. All protonated amines give rise to this readily observable absorption caused by the ammonium  $R_3N-H^+$  bond. Figure 25.12 gives an example.

## NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

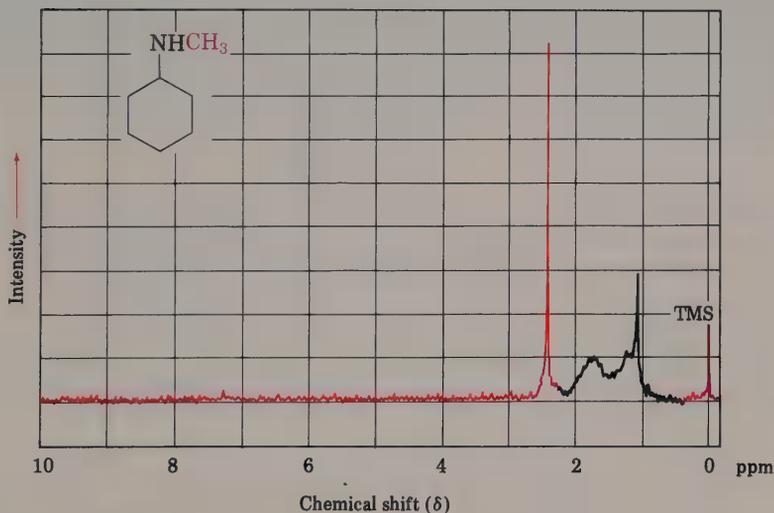
Amines are often difficult to identify solely by  $^1H$  NMR spectroscopy because N-H protons tend to appear as broad signals without clear-cut coupling to neighboring C-H protons. The situation is similar to that for hydroxyl protons (Section 17.11). As with hydroxyl O-H protons, amine N-H proton absorptions can appear over a wide range and are best identified by adding a small amount of  $D_2O$  to the sample tube. Exchange of N-D for N-H occurs, and the N-H signal disappears from the NMR spectrum.





**Figure 25.12** Infrared spectrum of triethylammonium chloride.

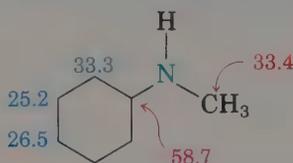
Protons on the carbon next to nitrogen are somewhat deshielded because of the electron-withdrawing effect of the nitrogen, and they therefore absorb at lower field than alkane protons do. *N*-Methyl groups are particularly distinctive, since they absorb as a sharp three-proton singlet at 2.2–2.6  $\delta$ . The *N*-methyl resonance at 2.42  $\delta$  is easily seen in the  $^1\text{H}$  NMR spectrum of *N*-methylcyclohexylamine (Figure 25.13).



**Figure 25.13** Proton NMR spectrum of *N*-methylcyclohexylamine.

Carbons next to amine nitrogens are slightly deshielded in the  $^{13}\text{C}$  NMR spectrum and absorb about 20 ppm downfield from where they would absorb in an alkane of similar structure. In *N*-methylcyclohexylamine, for example,

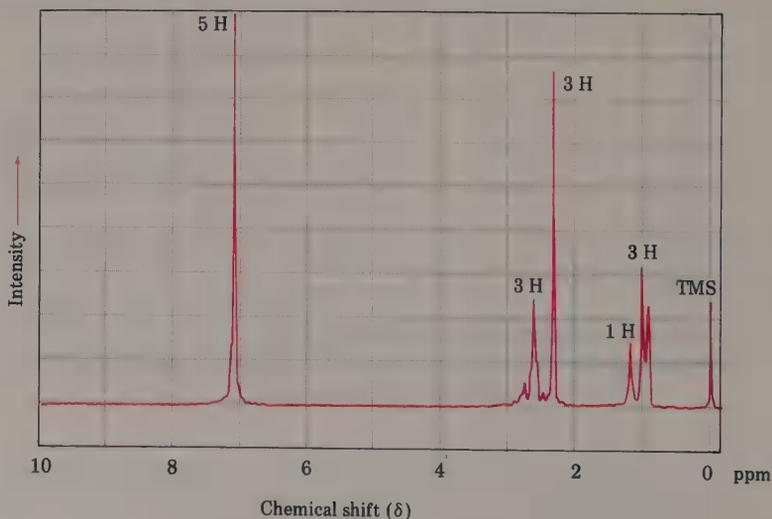
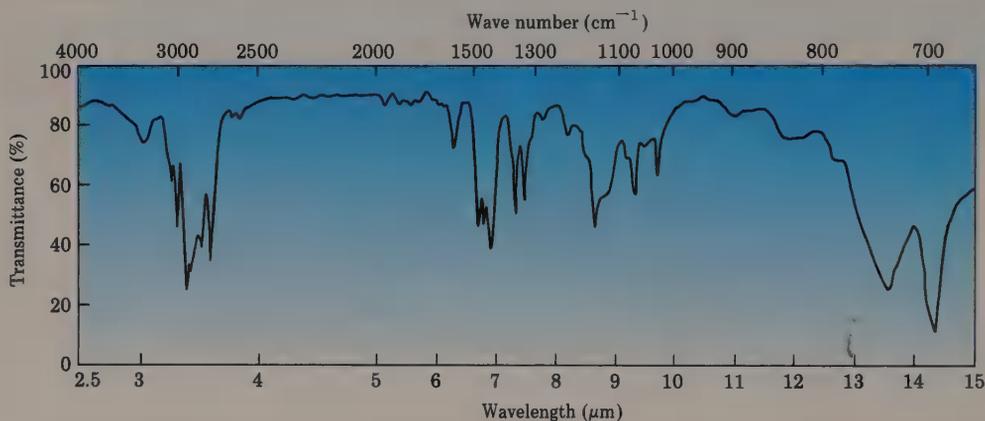
the ring carbon to which nitrogen is attached absorbs at a position 24 ppm lower than that of any other ring carbon (Figure 25.14).



**Figure 25.14**  $^{13}\text{C}$  NMR absorptions for *N*-methylcyclohexylamine.

**PROBLEM** .....

- 25.23** The infrared and  $^1\text{H}$  NMR spectra shown are those of methamphetamine, which has a molecular ion in the mass spectrum at  $m/z = 149$ . Propose a structure for methamphetamine, and justify your answer. [Note: The signal at 1.2  $\delta$  disappears when  $\text{D}_2\text{O}$  is added to the sample.]



## 25.12 Summary and Key Words

**Amines** are organic derivatives of ammonia. They are named in the IUPAC system either by adding the suffix *-amine* to the names of the alkyl substituents or by considering the amino group as a substituent on a more complex parent molecule.

Bonding in amines is similar to that in ammonia. The nitrogen atom is  $sp^3$  hybridized, the three substituents are directed to three corners of a tetrahedron, and the lone pair of nonbonding electrons occupies the fourth corner of the tetrahedron. An interesting feature of this tetrahedral structure is that amines undergo a rapid, umbrella-like pyramidal inversion, which interconverts mirror-image structures.

The chemistry of amines is dominated by the lone-pair electrons on nitrogen. Thus, amines are both basic and nucleophilic. The simplest method of amine synthesis involves  $S_N2$  reaction of ammonia or an amine with an alkyl halide. Alkylation of ammonia yields a primary amine; alkylation of a primary amine yields a secondary amine; and so on. This method often gives poor yields, however, and an alternative such as the **Gabriel amine synthesis** is often preferred.

Amines can also be prepared by a number of reductive methods. For example,  $LiAlH_4$  reduction of amides, nitriles, and azides yields amines. Even more important is the **reductive amination** reaction, whereby a ketone or an aldehyde is treated with an amine in the presence of a reducing agent such as  $NaBH_3CN$ . An intermediate imine is formed and then immediately reduced by the hydride reagent present.

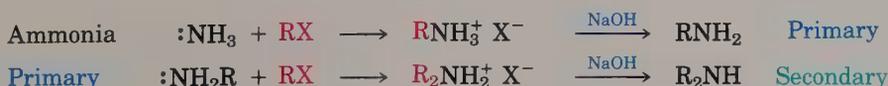
A final method of amine synthesis involves the **Hofmann and Curtius rearrangements** of carboxylic acid derivatives. Both methods involve migration of the  $R-$  group bonded to the carbonyl carbon and provide a product that has one less carbon atom than the starting material has.

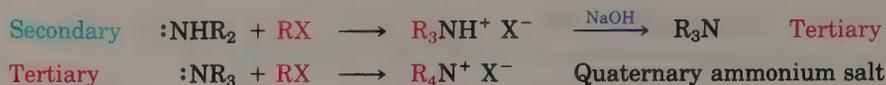
Many of the reactions of amines are familiar from past chapters. Thus, amines react with alkyl halides in  $S_N2$  reactions and with acid chlorides in nucleophilic acyl substitution reactions. Amines also undergo  $E2$  elimination to yield alkenes if they are first quaternized by treatment with methyl iodide and then heated with silver oxide (the **Hofmann elimination**).

Amines show a number of characteristic features that aid their spectroscopic identification. In the infrared spectrum,  $N-H$  absorptions are readily detectable at  $3200-3500\text{ cm}^{-1}$ . In the  $^1H$  NMR spectrum, protons on carbon next to nitrogen are deshielded and absorb in the range  $2.2-2.6\ \delta$ . In the mass spectrum, amines undergo a characteristic alpha cleavage.

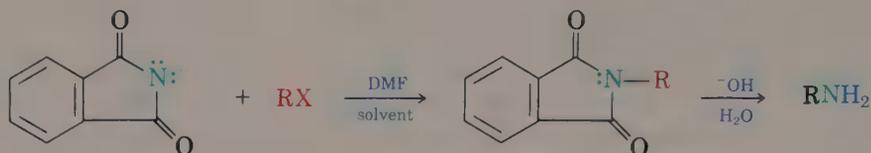
## 25.13 Summary of Reactions

1. Preparation of amines (Section 25.7)
  - a. The  $S_N2$  alkylation of alkyl halides

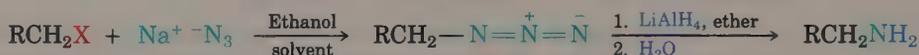




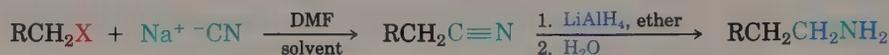
## b. Gabriel amine synthesis



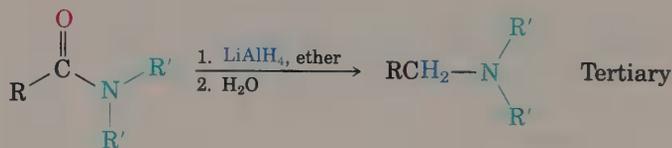
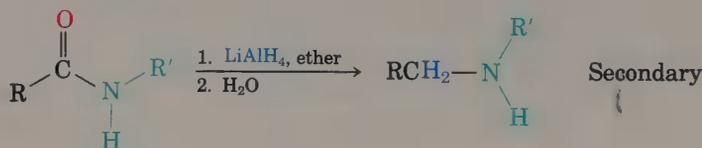
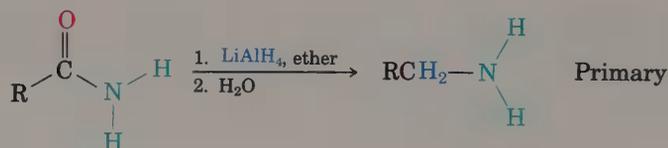
## c. Reduction of azides



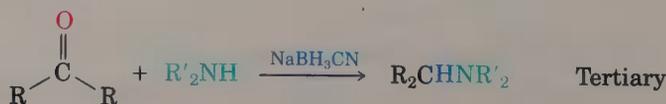
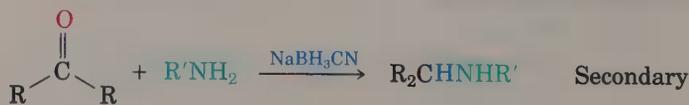
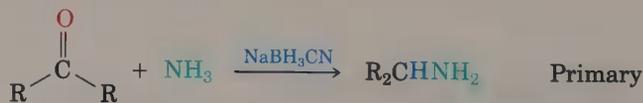
## d. Reduction of nitriles



## e. Reduction of amides



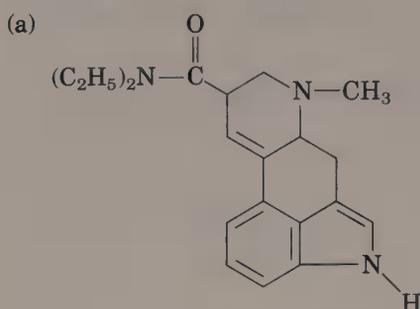
## f. Reductive amination of ketones/aldehydes



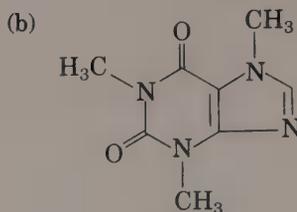


## ADDITIONAL PROBLEMS .....

25.24 Classify each of the amine nitrogen atoms in these substances as primary, secondary, or tertiary:



Lysergic acid diethylamide

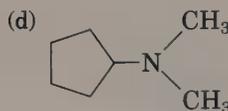
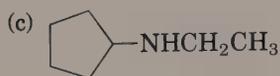
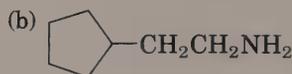
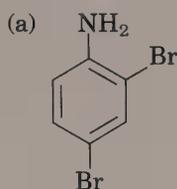


Caffeine

25.25 Draw structures corresponding to these IUPAC names:

- (a) *N,N*-Dimethylaniline (b) (Cyclohexylmethyl)amine  
 (c) *N*-Methylcyclohexylamine (d) (2-Methylcyclohexyl)amine  
 (e) 3-(*N,N*-Dimethylamino)propanoic acid  
 (f) *N*-Isopropyl-*N*-methylcyclohexylamine

25.26 Name these compounds by IUPAC rules:



25.27 How can you explain the fact that trimethylamine (bp 3°C) boils lower than dimethylamine (bp 7°C) even though it has a higher molecular weight?

25.28 How would you prepare these substances from 1-butanol?

- (a) Butylamine (b) Dibutylamine  
 (c) Propylamine (d) Pentylamine  
 (e) *N,N*-Dimethylbutylamine (f) Propene

25.29 How would you prepare these substances from pentanoic acid?

- (a) Pentanamide (b) Butylamine  
 (c) Pentylamine (d) 2-Bromopentanoic acid  
 (e) Hexanenitrile (f) Hexylamine

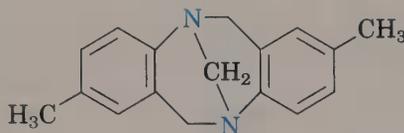
25.30 Treatment of bromoacetone with ammonia yields a compound having the formula  $\text{C}_6\text{H}_{10}\text{N}_2$ , rather than the expected 1-amino-2-propanone. What is a likely structure for the product?

25.31 Propose structures for substances that fit these descriptions:

- (a) A chiral quaternary ammonium salt  
 (b) A five-membered heterocyclic amine  
 (c) A secondary amine,  $\text{C}_6\text{H}_{11}\text{N}$

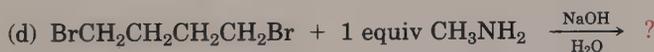
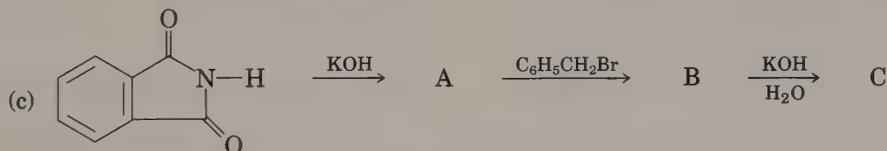
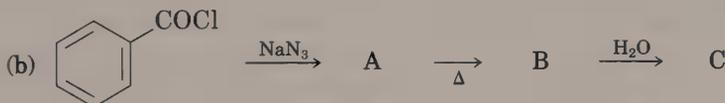
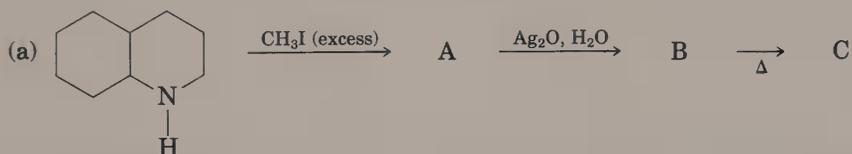
- 25.32 How might you prepare pentylamine from these starting materials?  
 (a) Pentanamide (b) Pentanenitrile (c) 1-Butene  
 (d) Hexanamide (e) 1-Butanol (f) 5-Decene  
 (g) Pentanoic acid

- 25.33 Although most chiral trisubstituted amines cannot be resolved into enantiomers because nitrogen pyramidal inversion occurs too rapidly, one exception to this generalization is the substance known as Tröger's base. Make molecular models of Tröger's base, and then explain why it is resolvable into enantiomers.



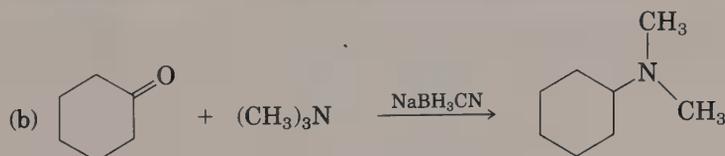
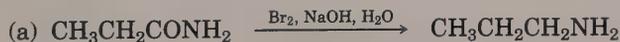
Tröger's base

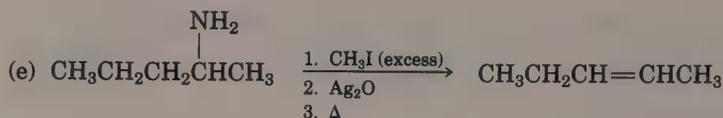
- 25.34 Predict the product(s) of these reactions. If more than one product is formed, tell which is major.



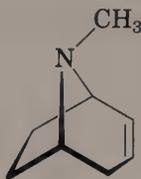
- 25.35 Phthalimide used in the Gabriel synthesis is prepared by reaction of ammonia with phthalic anhydride (1,2-benzenedicarboxylic anhydride). Propose a mechanism for the reaction.

- 25.36 The following syntheses are incorrect. What is wrong with each?



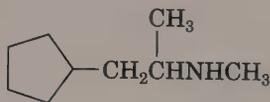


- 25.37 Coniine,  $\text{C}_8\text{H}_{17}\text{N}$ , is the toxic principle of poison hemlock drunk by Socrates. When subjected to Hofmann elimination, coniine yields 5-(*N,N*-dimethylamino)-1-octene. When subjected to the Hinsberg test, coniine yields a benzenesulfonamide derivative that is insoluble in base. What is the structure of coniine?
- 25.38 Atropine,  $\text{C}_{17}\text{H}_{23}\text{NO}_3$ , is a poisonous alkaloid isolated from the leaves and roots of *Atropa belladonna*, the deadly nightshade. In low doses, atropine acts as a muscle relaxant; 0.5 ng (nanogram;  $10^{-9}$  g) is sufficient to cause pupil dilation. On basic hydrolysis, atropine yields tropic acid,  $\text{C}_6\text{H}_5\text{CH}(\text{CH}_2\text{OH})\text{COOH}$ , and tropine,  $\text{C}_8\text{H}_{15}\text{NO}$ . Tropine is an optically inactive alcohol that yields tropidene on dehydration with  $\text{H}_2\text{SO}_4$ . Propose a structure for atropine.



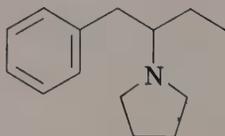
Tropidene

- 25.39 Tropidene (Problem 25.38) can be converted by a series of steps into tropilidene (1,3,5-cycloheptatriene). How would you accomplish this conversion?
- 25.40 One problem with reductive amination as a method of amine synthesis is that by-products are sometimes obtained. For example, reductive amination of benzaldehyde with methylamine leads to a mixture of methylbenzylamine and methyl-dibenzylamine. How do you suppose the tertiary amine by-product is formed? Propose a mechanism.
- 25.41 What are the major products you would expect from Hofmann elimination of these amines?
- (a) *N*-Methylcyclopentylamine
- (b)  $(\text{CH}_3)_2\text{CHCH}(\text{NH}_2)\text{CH}_2\text{CH}_2\text{CH}_3$
- (c) *N*-Phenyl-*N*-(2-hexyl)amine
- 25.42 Cyclopentamine is an amphetamine-like central nervous system stimulant. Propose a synthesis of cyclopentamine from materials of five carbons or less.



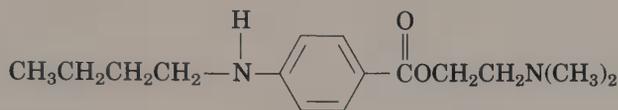
Cyclopentamine

- 25.43 Prolitane is an antidepressant drug that is prepared commercially using a reductive amination. What amine and what carbonyl precursors are used?



Prolitane

- 25.44 Tetracaine is a substance used medicinally as a spinal anesthetic during lumbar punctures (spinal taps).

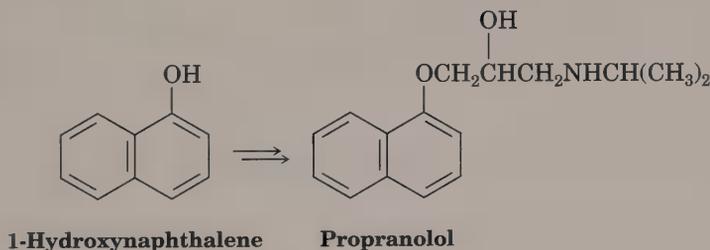


Tetracaine

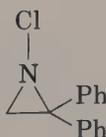
- (a) How would you prepare tetracaine from the corresponding aniline derivative,  $\text{ArNH}_2$ ?  
 (b) How would you prepare tetracaine from *p*-nitrobenzoic acid?  
 (c) How would you prepare tetracaine from benzene?
- 25.45 Propose a structure for the product of formula  $\text{C}_9\text{H}_{17}\text{N}$  that results when 2-(2-cyanoethyl)cyclohexanone is reduced catalytically.



- 25.46 How would you synthesize coniine (Problem 25.37) from acrylonitrile ( $\text{H}_2\text{C}=\text{CHCN}$ ) and ethyl 3-oxohexanoate ( $\text{CH}_3\text{CH}_2\text{CH}_2\text{COCH}_2\text{CO}_2\text{C}_2\text{H}_5$ )? [Hint: See Problem 25.45.]
- 25.47 How would you synthesize the heart stimulant propranolol starting from 1-hydroxynaphthalene and any other reagents needed?



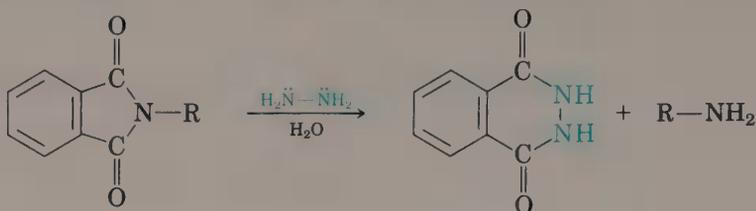
- 25.48 Although the barrier to nitrogen inversion is normally too low to allow isolation of one enantiomer of a chiral amine, (+)-1-chloro-2,2-diphenylaziridine has been prepared optically pure and has been shown to be stable for several hours at  $0^\circ\text{C}$ . Propose an explanation of this unusual stability.



1-Chloro-2,2-diphenylaziridine

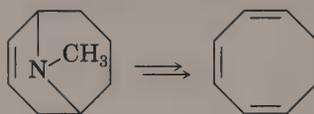
- 25.49 The hydrolysis of phthalimides is often slow, and an alternative method is sometimes needed to liberate the primary amine in the last stage of a Gabriel synthesis. In the

Ing–Manske modification, reaction of an *N*-alkylphthalimide with hydrazine is employed. Propose a mechanism for this *hydrazinolysis*.



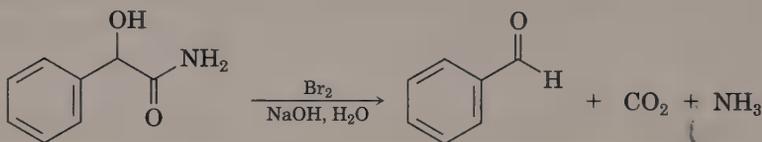
Phthalhydrazide

- 25.50 Cyclooctatetraene was first synthesized in 1911 by a route that involved the following transformation:

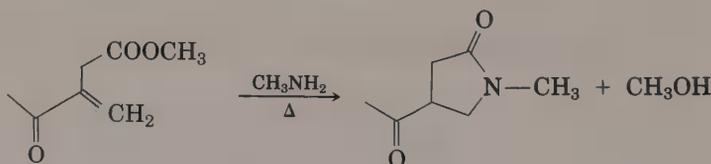


How might you use the Hofmann elimination to accomplish this reaction? How would you finish the synthesis by converting cyclooctatriene into cyclooctatetraene?

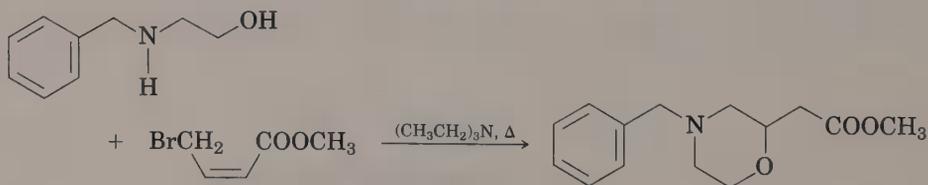
- 25.51 When an  $\alpha$ -hydroxy amide is treated with  $\text{Br}_2$  in aqueous  $\text{NaOH}$  under Hofmann rearrangement conditions, loss of  $\text{CO}_2$  occurs and a chain-shortened aldehyde is formed. Propose a mechanism for the reaction.



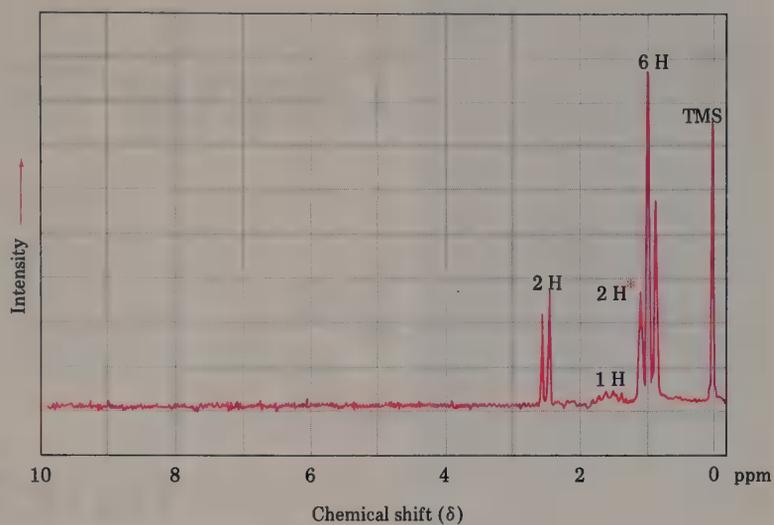
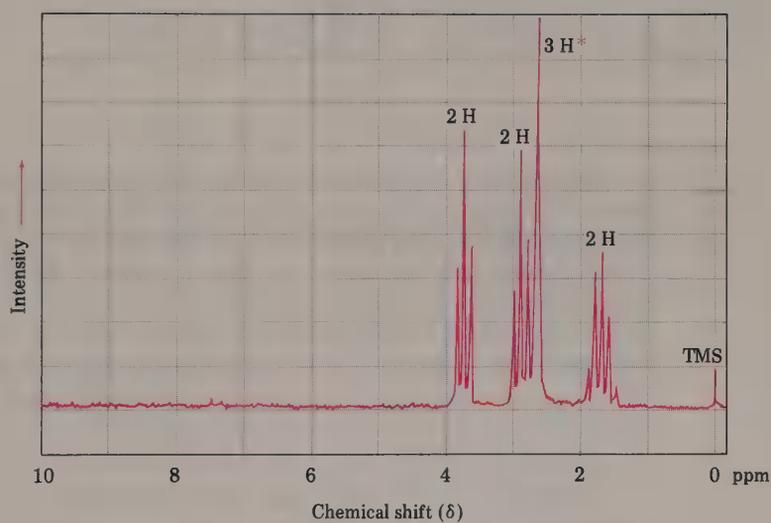
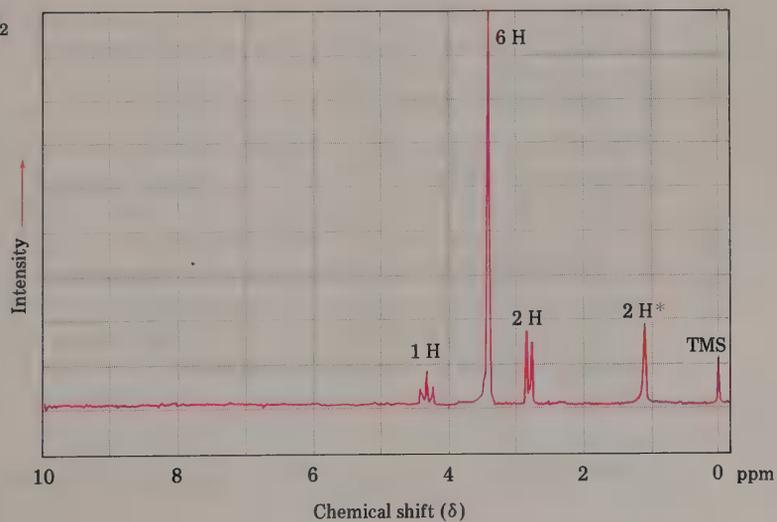
- 25.52 Propose a mechanism for the following reaction:

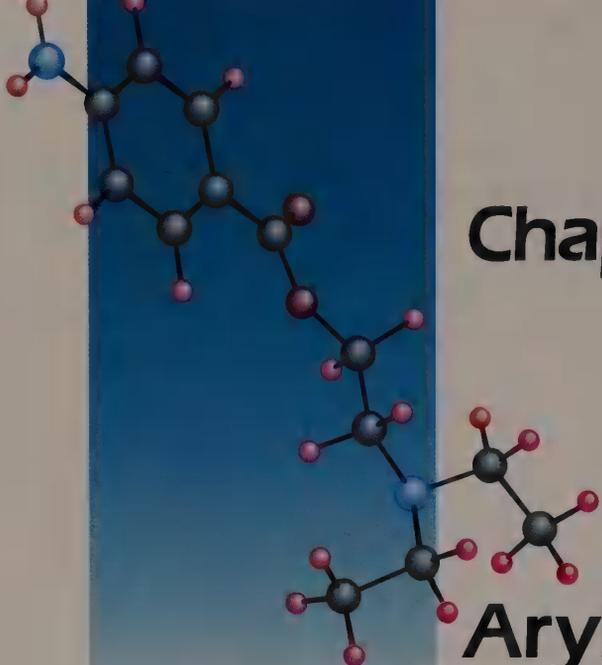


- 25.53 Propose a mechanism for the following reaction:



- 25.54 Propose structures for amines with the following  $^1\text{H}$  NMR spectra. The peaks marked by an asterisk disappear when  $\text{D}_2\text{O}$  is added to the sample.

(a)  $C_4H_{11}N$ 

 (b)  $C_3H_9NO$ 

 (c)  $C_4H_{11}NO_2$ 




## Chapter 26

# Arylamines and Phenols

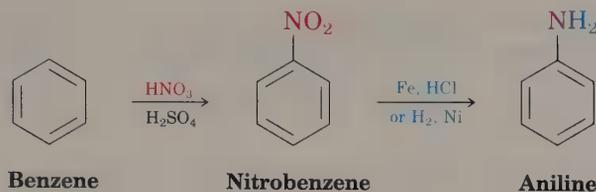
The founding of the modern organic chemical industry can be traced to the need for a single organic compound—aniline—and to the activities of one person—Sir William Henry Perkin.<sup>1</sup> Perkin, a student of Hofmann's at the Royal College of Chemistry in London, worked during the day on problems assigned him by Hofmann but spent his free time working on his own ideas in an improvised home laboratory. One day during Easter vacation in 1856, he decided to examine the oxidation of aniline with potassium dichromate. Although the reaction appeared unpromising at first, yielding a tarry black product, Perkin was able by careful extraction with methanol to isolate a small amount of a beautiful purple pigment that had the properties of a dye.

Since the only dyes known at the time were the naturally occurring vegetable dyes such as indigo, Perkin's synthetic purple dye, which he named *mauve*, created a sensation. Realizing the possibilities, Perkin resigned his post with Hofmann and, at the age of 18, formed a company to exploit his remarkable discovery.

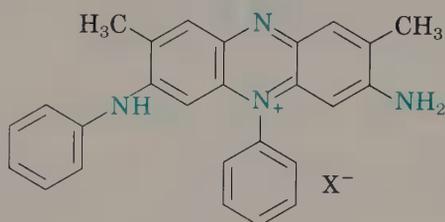
No chemical industry existed at the time, since there had never before been a need for synthetic chemicals. Large-scale chemical manufacture was unknown, and Perkin's first task was to devise a procedure for preparing the needed quantities of aniline. He therefore worked out the techniques of manufacture and soon learned to prepare aniline on a large scale by nitration of benzene, followed by reduction of nitrobenzene with iron and hydrochloric acid. A similar procedure is used today to prepare some 500,000 tons of aniline annually in the United States, although the reduction step is now carried out by catalytic hydrogenation.

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<sup>1</sup>Sir William Henry Perkin (1838–1907); b. London; studied at Royal College of Chemistry, London; industrial consultant, London.



Subsequent work showed that Perkin's original mauve was not derived from aniline but from a toluidine (methylaniline) impurity in his starting material. Pure aniline yields a similar dye, however, which came to be marketed under the name *pseudomauveine*.



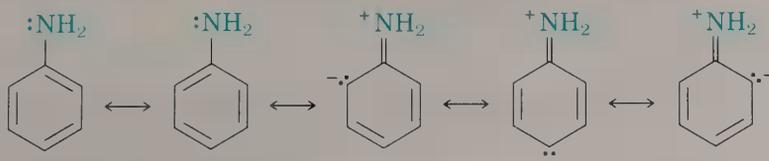
Perkin's mauve  
(pseudomauveine has no methyl groups)

Dyestuff manufacture is today a thriving and important part of the chemical industry, and many commonly used pigments are derived from aniline. Although aniline itself and several substituted anilines are available naturally from coal tar, synthesis from benzene is the major source.

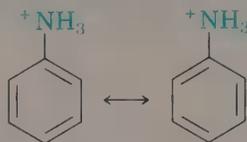
## 26.1 Basicity of Arylamines

**Arylamines**, like their aliphatic counterparts, are basic; the lone pair of nonbonding electrons on nitrogen can bond to Lewis acids, yielding an arylammonium salt. The base strength of arylamines is generally lower than that of aliphatic amines, however. Thus, methylammonium ion has  $\text{p}K_{\text{a}} = 10.64$ , whereas anilinium ion has  $\text{p}K_{\text{a}} = 4.63$ . [Remember: The base strength of an amine is inversely related to the acid strength of its corresponding ammonium ion (Section 25.4). A stronger base like methylamine corresponds to a less acidic ammonium ion (higher  $\text{p}K_{\text{a}}$ ), whereas a weaker base like aniline corresponds to a more acidic ammonium ion (lower  $\text{p}K_{\text{a}}$ ).]

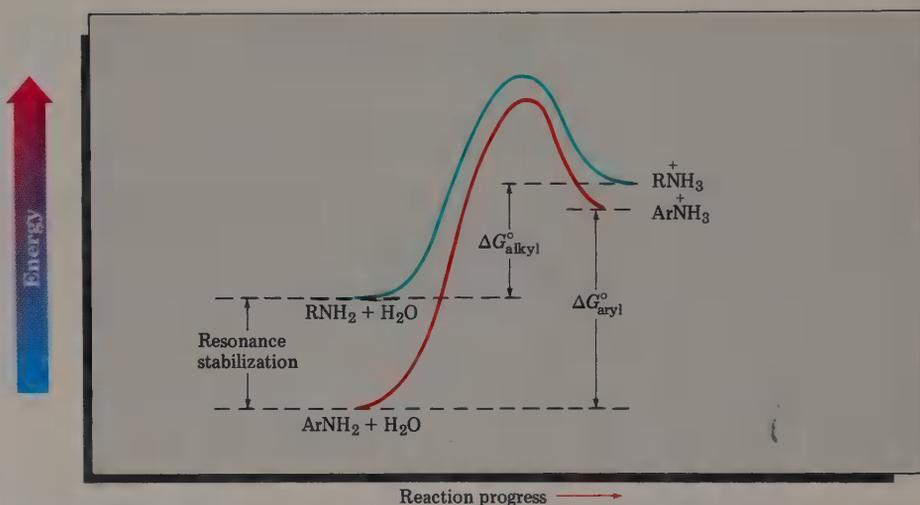
Arylamines are less basic than alkylamines because the nitrogen lone-pair electrons are delocalized by orbital overlap with the aromatic ring pi electron system and are less available for bonding. In resonance terms, arylamines are stabilized relative to alkylamines because of the five contributing resonance structures:



Resonance stabilization is lost on protonation, though, since only two resonance structures are possible for the arylammonium ion:



As a result, the energy difference ( $\Delta G^\circ$ ) between protonated and nonprotonated forms is higher for arylamines than it is for alkylamines, and arylamines are therefore less basic. Figure 26.1, which compares reaction energy diagrams for protonation of alkylamines and arylamines, illustrates the difference in  $\Delta G^\circ$  for the two reactions.



**Figure 26.1** Reaction energy diagrams for the protonation of alkylamines (green curve) and arylamines (red curve). Arylamines have a larger  $\Delta G^\circ$  and are therefore less basic than alkylamines, primarily because of resonance stabilization of their ground state.

Substituted arylamines can be either more basic or less basic than aniline, depending on the nature of the substituent. Table 26.1 presents data for a variety of para-substituted anilines.

As a general rule, substituents that increase the reactivity of an aromatic ring toward electrophilic substitution ( $-\text{CH}_3$ ,  $-\text{NH}_2$ ,  $-\text{OCH}_3$ ) also increase the basicity of the corresponding arylamine. Conversely, substituents that decrease ring reactivity ( $-\text{Cl}$ ,  $-\text{NO}_2$ ,  $-\text{CN}$ ) also decrease arylamine basicity. Although Table 26.1 considers only para-substituted anilines, the same general trends are observed for ortho and meta derivatives.

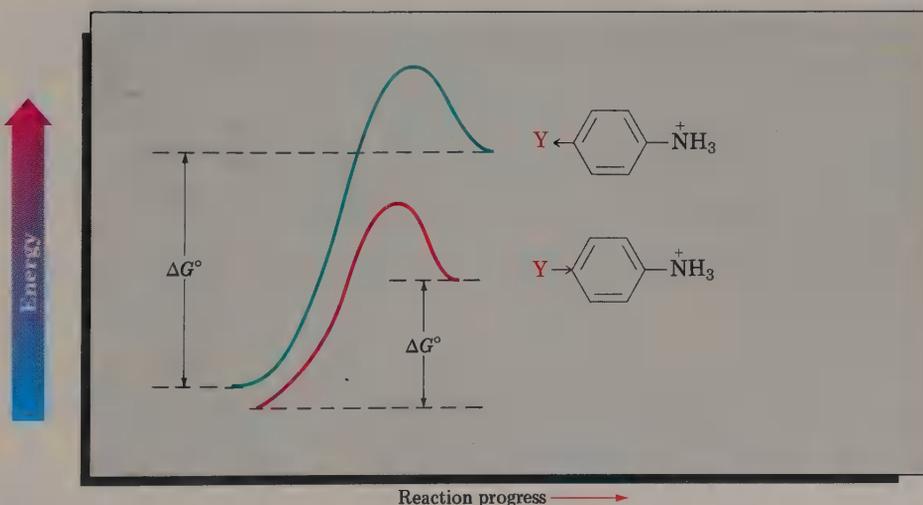
The best way to understand the effect of substituent groups on arylamine basicity is to look at reaction energy diagrams for the amine protonation step (Figure 26.2). Activating substituents make the aromatic ring

TABLE 26.1 Base Strength of Para-substituted Anilines



	Substituent, Y	pK <sub>a</sub>	
Stronger base 	—NH <sub>2</sub>	6.15	Activating groups
	—OCH <sub>3</sub>	5.34	
	—CH <sub>3</sub>	5.08	
Weaker base	—H	4.63	Deactivating groups
	—Cl	3.98	
	—Br	3.86	
	—C≡N	1.74	
	—NO <sub>2</sub>	1.00	

electron-rich, thereby increasing the stability of the positively charged ammonium ion. Deactivating substituents, however, make the aromatic ring electron-poor, thereby decreasing the stability of the positively charged ion. We therefore find a lower  $\Delta G^\circ$  for protonation of an activated arylamine than for protonation of a deactivated arylamine.



**Figure 26.2** Reaction energy diagrams for the protonation of substituted arylamines. An electron-donating substituent (red curve) stabilizes the ammonium salt much more than an electron-withdrawing substituent does (green curve).

## PROBLEM .....

- 26.1 How do you account for the fact that *p*-nitroaniline ( $pK_a = 1.0$ ) is less basic than *m*-nitroaniline ( $pK_a = 2.5$ ) by a factor of 30? Draw resonance structures to support your answer. (The  $pK_a$  values refer to the corresponding ammonium ions.)

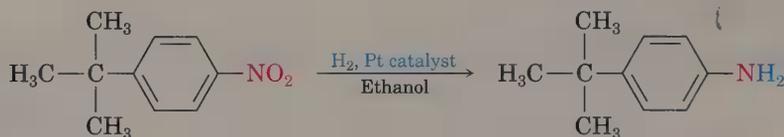
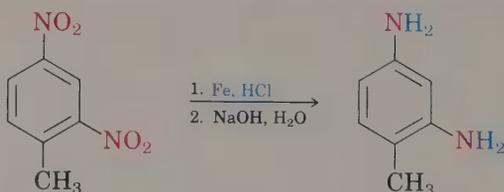
## PROBLEM .....

- 26.2 Rank the following compounds in order of ascending basicity. Don't look at Table 26.1 to decide your answers.
- (a) *p*-Nitroaniline, *p*-aminobenzaldehyde, *p*-bromoaniline  
 (b) *p*-Chloroaniline, *p*-aminoacetophenone, *p*-methylaniline  
 (c) *p*-(Trifluoromethyl)aniline, *p*-methylaniline, *p*-(fluoromethyl)aniline

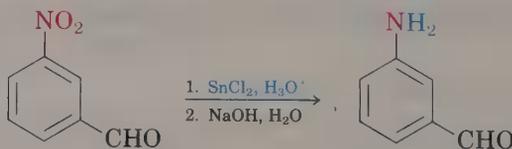
## 26.2 Preparation of Arylamines

Arylamines are usually prepared by nitration of an aromatic starting material, followed by reduction of the nitro group. No other method of synthesis approaches this nitration–reduction route for versatility and generality.

The reduction step can be carried out in many different ways, depending on the circumstances. Catalytic hydrogenation over platinum is clean and gives high yields, but is often incompatible with the presence of other reducible groups such as carbon–carbon double bonds or carbonyl groups elsewhere in the molecule. Iron, zinc, tin metal, and stannous chloride ( $\text{SnCl}_2$ ) are also effective when used in acidic aqueous solution. Stannous chloride is particularly mild and is often used when other reducible functional groups are present.

*p*-*tert*-Butylnitrobenzene*p*-*tert*-Butylaniline (100%)

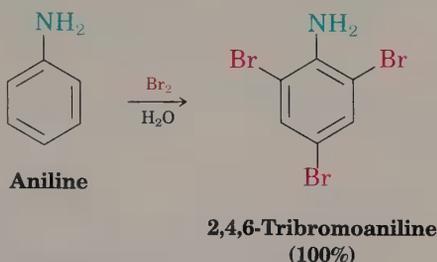
2,4-Dinitrotoluene

Toluene-2,4-diamine  
(74%)*m*-Nitrobenzaldehyde*m*-Aminobenzaldehyde  
(90%)

## 26.3 Reactions of Arylamines

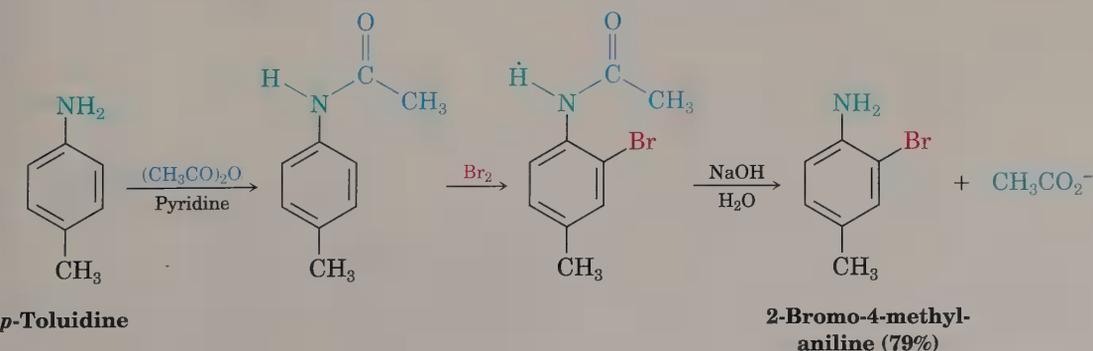
### ELECTROPHILIC AROMATIC SUBSTITUTION

Amino substituents are strongly activating, ortho- and para-directing groups in electrophilic aromatic substitution reactions (Section 16.5). The high reactivity of amino-substituted benzenes can be a drawback at times, since it's sometimes difficult to prevent polysubstitution. For example, reaction of aniline with bromine takes place rapidly to yield the 2,4,6-tribrominated product. The amino group is so strongly activating that it's not possible to stop cleanly at the monobromo stage.

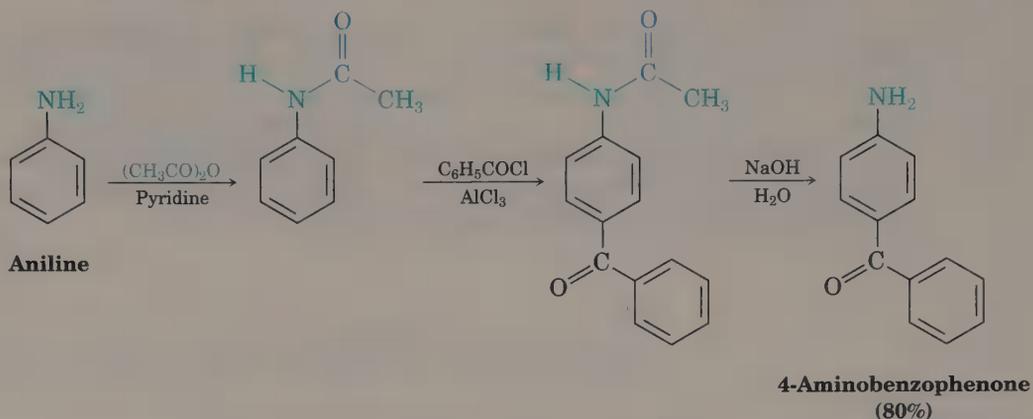


Another drawback to the use of amino-substituted benzenes in electrophilic aromatic substitution reactions is that Friedel–Crafts reactions are not successful (Section 16.3). The amino group forms an acid–base complex with the aluminum chloride Friedel–Crafts catalyst, which prevents further reaction from occurring. Both of these drawbacks, high reactivity and amine basicity, can be overcome by carrying out electrophilic aromatic substitution reactions on the corresponding amide, rather than on the free amine.

As we saw in Section 21.6, treatment of an arylamine with acetic anhydride yields an *N*-acetylated product. Though still ortho- and para-directing and activating, amido substituents (–NHCOR) are much less strongly activating and much less basic than amino groups because their nitrogen lone-pair electrons are delocalized by overlap with the neighboring carbonyl group (Section 25.4). As a result, bromination of an *N*-arylamide occurs cleanly to give a monobromo product, which can be hydrolyzed by aqueous base to give the free bromoamine. For example, *p*-toluidine (4-methylaniline) can be acetylated, brominated, and hydrolyzed to yield 2-bromo-4-methylaniline in 79% yield. None of the 2,6-dibrominated product is obtained.

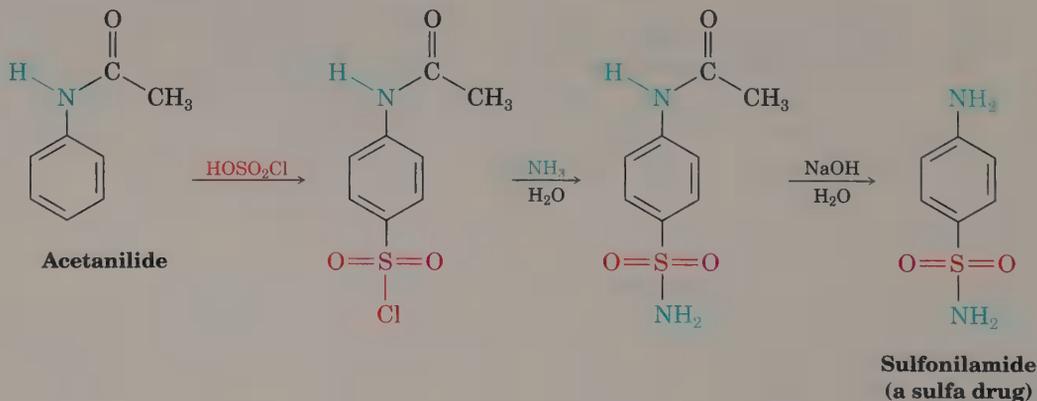


Friedel–Crafts alkylations and acylations of *N*-arylamides also proceed normally. For example, benzoylation of acetanilide (*N*-acetylaniline) under Friedel–Crafts conditions gives 4-aminobenzophenone in 80% yield after hydrolysis:



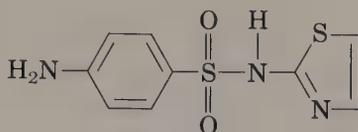
The simple chemical trick of modulating the reactivity of amino-substituted benzenes by forming an amide is an extremely useful one that allows many kinds of electrophilic aromatic substitutions to be carried out that would otherwise be impossible. A good example is the preparation of *sulfa* drugs.

Sulfa drugs such as sulfanilamide were among the first pharmaceutical agents to be used clinically against infection. Although they have largely been replaced today by safer and more powerful antibiotics, sulfa drugs were widely used in the 1940s and were credited with saving the lives of thousands of wounded during World War II. They are prepared by chlorosulfonation of acetanilide, followed by reaction of *p*-(*N*-acetylamino)benzenesulfonyl chloride with ammonia or some other amine to give a sulfonamide. Hydrolysis of the amide then yields the sulfa drug. Note that this hydrolysis can be carried out in the presence of the sulfonamide group, because sulfonamides hydrolyze very slowly.



PROBLEM.....

**26.3** Propose a synthesis of sulfathiazole from benzene and any necessary amine.



Sulfathiazole

## PROBLEM

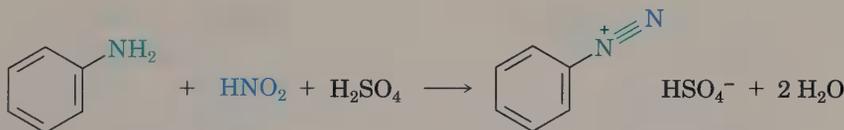
- 26.4 Account for the fact that an amido substituent ( $-\text{NHCOR}$ ) is ortho- and para-directing, by drawing resonance structures that share the nitrogen lone-pair electrons with the aromatic ring.

## PROBLEM

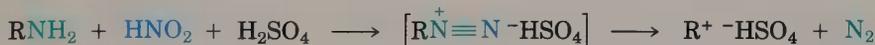
- 26.5 Propose syntheses of these compounds from benzene:
- (a) *N,N*-Dimethylaniline                      (b) *p*-Chloroaniline  
 (c) *m*-Chloroaniline                        (d) 2,4-Dimethylaniline

## DIAZONIUM SALTS: THE SANDMEYER REACTION

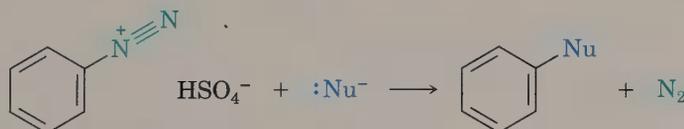
Primary aromatic amines react with nitrous acid,  $\text{HNO}_2$ , to yield stable **arenediazonium salts**,  $\text{Ar}-\overset{+}{\text{N}}\equiv\text{N} \text{X}^-$ . This **diazotization** reaction is compatible with the presence of a wide variety of substituents on the aromatic ring.



Alkylamines also react with nitrous acid, but the products of these reactions, alkanediazonium salts, are too reactive to isolate since they lose nitrogen instantly to yield carbocations.

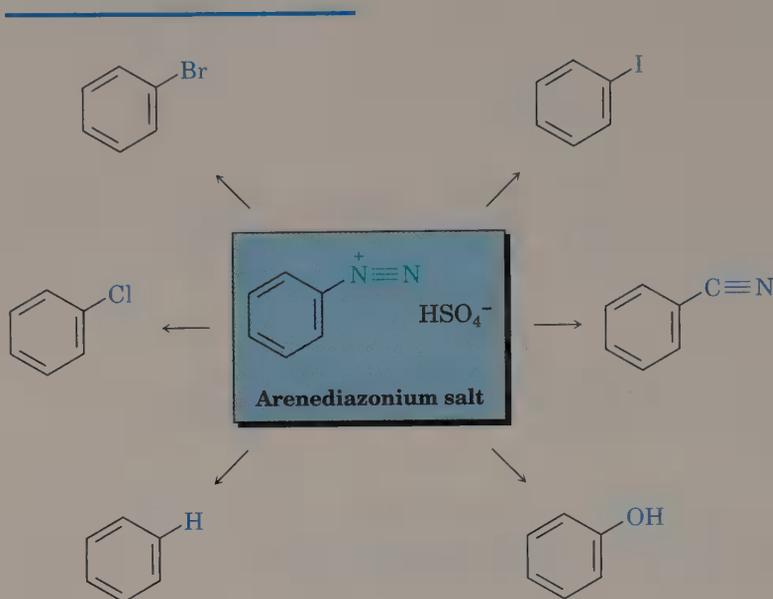


Arenediazonium salts are extremely useful in synthesis, because the diazonio group ( $\text{N}_2^+$ ) can be replaced by nucleophiles:



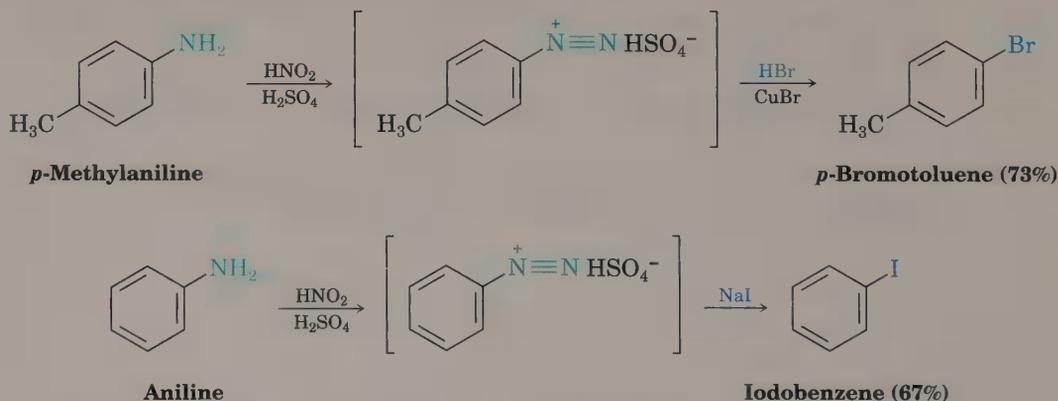
Many different nucleophiles react with arenediazonium salts, and many different substituted benzenes can be prepared with this reaction. The overall sequence of (1) nitration, (2) reduction, (3) diazotization, and (4) nucleophile replacement is probably the single most versatile method of aromatic substitution (Figure 26.3). Although the details aren't fully understood, the

mechanism by which these substitutions occur is a radical, rather than a polar, one.



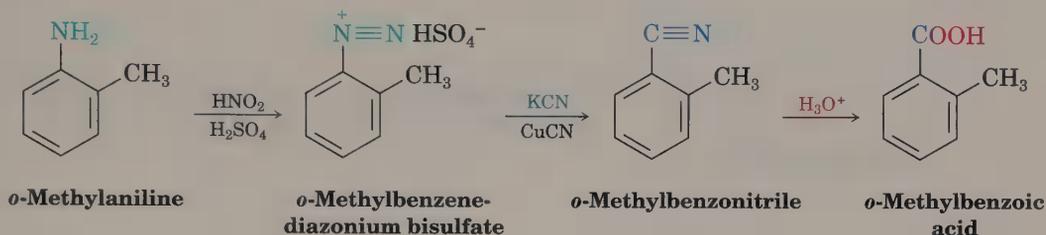
**Figure 26.3** Preparation of substituted aromatic compounds by diazonio replacement reactions.

Aryl chlorides and bromides are prepared by reaction of an arenediazonium salt with the corresponding cuprous halide,  $\text{CuX}$ , a process called the **Sandmeyer<sup>2</sup> reaction**. Aryl iodides can be prepared by direct reaction with sodium iodide without using a cuprous salt. Yields generally fall in the 60–80% range.

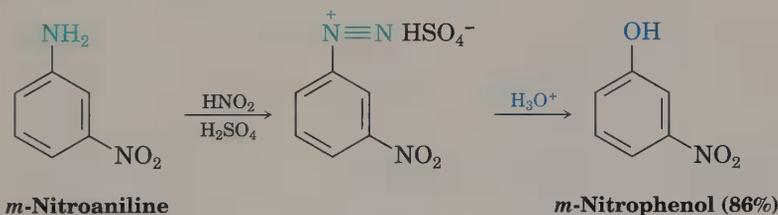


<sup>2</sup>Traugott Sandmeyer (1854–1922); b. Wettingen, Switzerland; Ph.D. Heidelberg (Gattermann); Geigy Company, Basel, Switzerland.

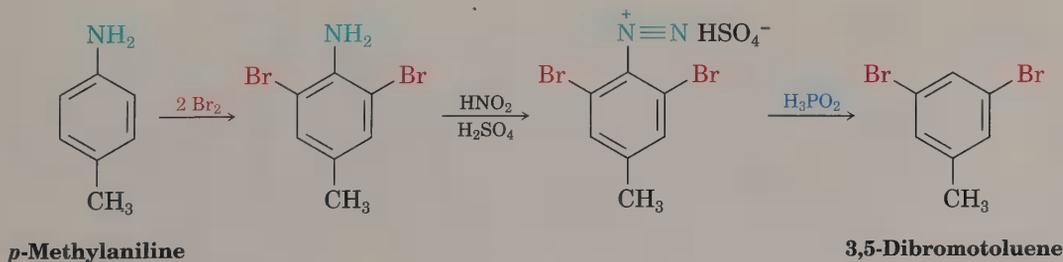
Similar treatment of an arenediazonium salt with cuprous cyanide yields the arenenitrile,  $\text{ArCN}$ . This reaction is particularly useful, because it allows the replacement of a nitrogen substituent by a carbon substituent. The nitrile can then be further converted into other functional groups such as carboxyl,  $-\text{COOH}$ . For example, hydrolysis of *o*-methylbenzonitrile, produced by Sandmeyer reaction of *o*-methylbenzenediazonium bisulfate with cuprous cyanide, yields *o*-methylbenzoic acid. This product could not be prepared from *o*-xylene by the usual side-chain oxidation route, since both methyl groups would be oxidized.

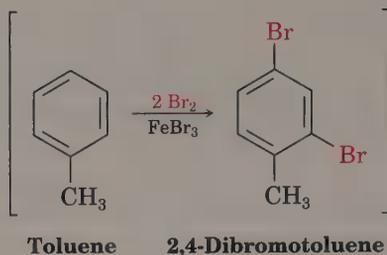


The diazonio group can also be replaced by  $-\text{OH}$  to yield phenols and by  $-\text{H}$  to yield arenes. Phenols are usually prepared by addition of the arenediazonium salt to hot aqueous acid. This reaction is especially important, since few other general methods exist for introducing an  $-\text{OH}$  group onto an aromatic ring.



Arenes are produced by reduction of the diazonium salt with hypophosphorous acid,  $\text{H}_3\text{PO}_2$ . This reaction is not particularly useful, however, unless there's a need for temporarily introducing an amino substituent onto a ring to take advantage of its activating effect. The preparation of 3,5-dibromotoluene from *p*-methylaniline (*p*-toluidine) illustrates how this can be done. Dibromination of *p*-toluidine occurs ortho to the strongly directing amino substituent, and diazotization followed by treatment with hypophosphorous acid yields 3,5-dibromotoluene. This product can't be prepared by direct bromination of toluene, however, since reaction would occur at positions 2 and 4.





## PROBLEM .....

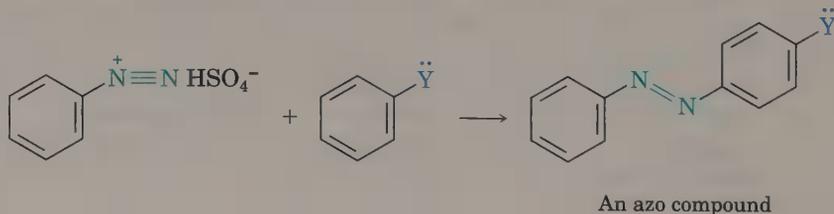
- 26.6 Alkanediazonium salts can't be isolated because they spontaneously lose  $N_2$  to yield a carbocation. Why do you suppose arenediazonium salts are more stable than alkanediazonium salts?

## PROBLEM .....

- 26.7 How would you prepare these compounds from benzene, using a diazonium replacement reaction at some point?
- |                                  |                                  |
|----------------------------------|----------------------------------|
| (a) <i>p</i> -Bromobenzoic acid  | (d) <i>p</i> -Methylbenzoic acid |
| (b) <i>m</i> -Bromobenzoic acid  | (e) 1,2,4-Tribromobenzene        |
| (c) <i>m</i> -Bromochlorobenzene |                                  |
- .....

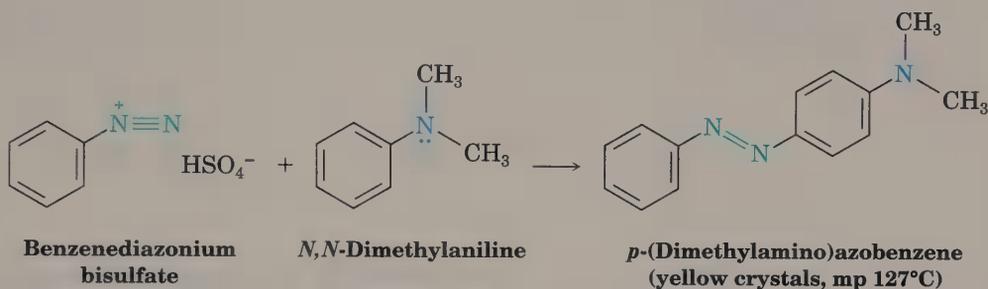
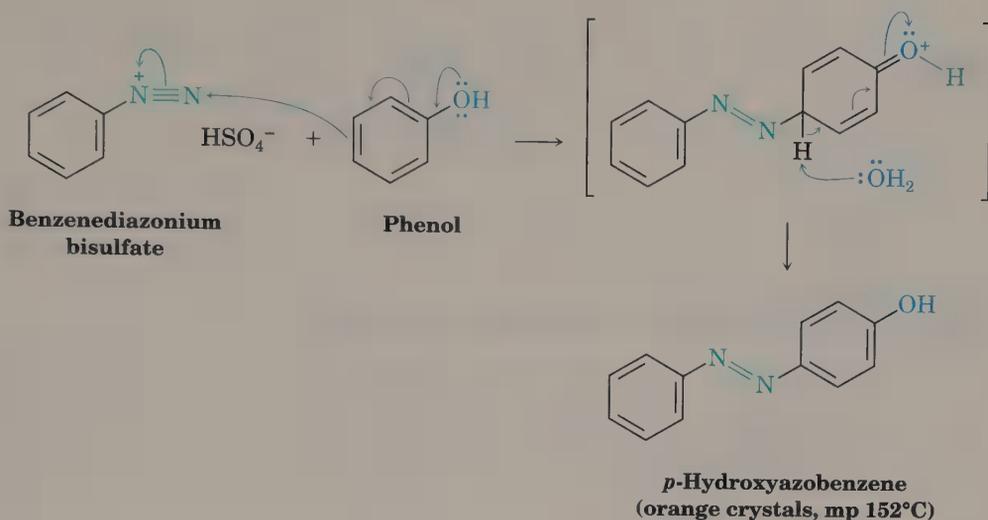
## DIAZONIUM COUPLING REACTIONS

In addition to their reactivity in Sandmeyer-type substitution reactions, arenediazonium salts undergo a coupling reaction with activated aromatic rings to yield brightly colored **azo compounds**,  $Ar-N=N-Ar'$ :

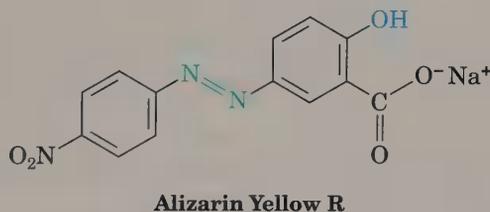


where  $Y = -OH$  or  $-NR_2$

Diazonium coupling reactions are typical electrophilic aromatic substitutions (Section 16.2) in which the positively charged diazonium ion is the electrophile that reacts with the electron-rich ring of a phenol or an arylamine. Reaction almost always occurs at the para position, although ortho attack can take place if the para position is blocked.



Azo-coupled products are widely used as dyes because their extended conjugated pi electron system causes them to absorb in the visible region of the electromagnetic spectrum. For example, *p*-(dimethylamino)azobenzene is bright yellow and was at one time used as a coloring agent in margarine. Another azo compound, Alizarin Yellow R, is used for dyeing wool.

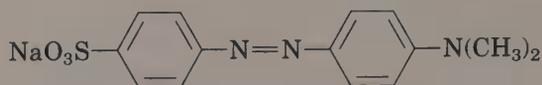


PROBLEM.....

26.8 Propose a synthesis of *p*-(dimethylamino)azobenzene from benzene.

PROBLEM.....

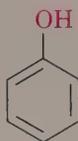
26.9 Methyl Orange is an azo dye that is widely used as a pH indicator. How would you synthesize Methyl Orange from benzene?



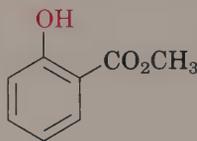
Methyl Orange

## 26.4 Phenols

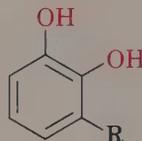
**Phenols** are compounds with a hydroxyl group directly bonded to an aromatic ring,  $\text{ArOH}$ . They occur widely throughout nature, and they serve as intermediates in the industrial synthesis of products as diverse as adhesives and antiseptics. Phenol itself is a general disinfectant found in coal tar; methyl salicylate is a flavoring agent and liniment found in oil of wintergreen; and the urushiols are the main allergenic constituents of poison oak and poison ivy. Note that the word *phenol* is the name both of a specific compound and of a class of compounds.



Phenol  
(also known as  
carbolic acid)



Methyl salicylate



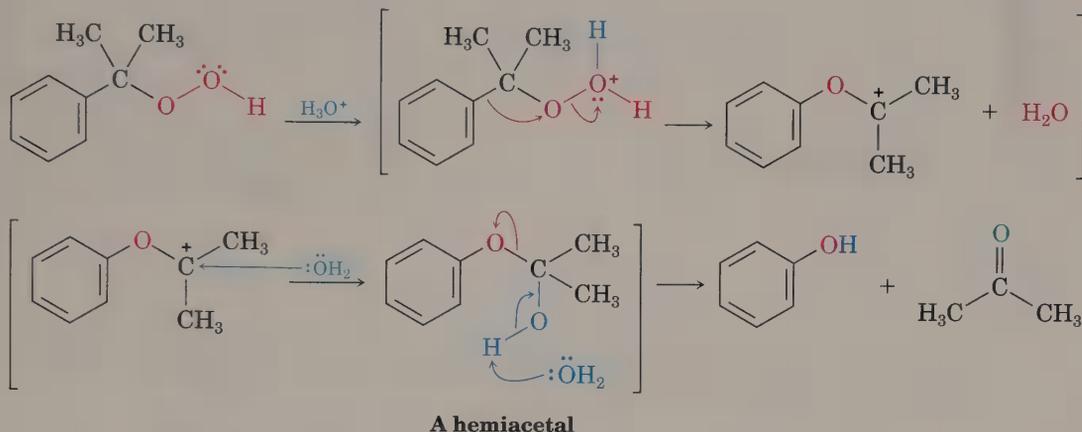
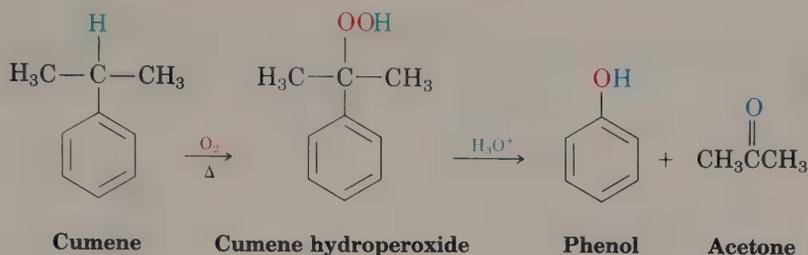
Urushiols  
(R = different  $\text{C}_{15}$  alkyl  
and alkenyl chains)

## 26.5 Industrial Uses of Phenols

The outbreak of World War I provided a stimulus for industrial preparation of large amounts of synthetic phenol, needed as raw material for the manufacture of the explosive, picric acid (2,4,6-trinitrophenol). Today, approximately 1.8 million tons of phenol per year are manufactured in the United States for use in a variety of products, including Bakelite resin and adhesives for binding plywood.

For many years, phenol was manufactured by the Dow process in which chlorobenzene reacts with sodium hydroxide at high temperature and pressure (Section 16.10). Now, however, an alternative synthesis from isopropylbenzene (cumene) is used. Cumene reacts with air at high temperature by a radical mechanism to form cumene hydroperoxide, which is converted into phenol and acetone by treatment with acid. This is a particularly efficient process, since two valuable chemicals are prepared at the same time.

The reaction occurs by protonation of oxygen, followed by rearrangement of the phenyl group from carbon to oxygen, and concurrent loss of water. Readdition of water then yields a hemiacetal, which breaks down to phenol and acetone (Figure 26.4).

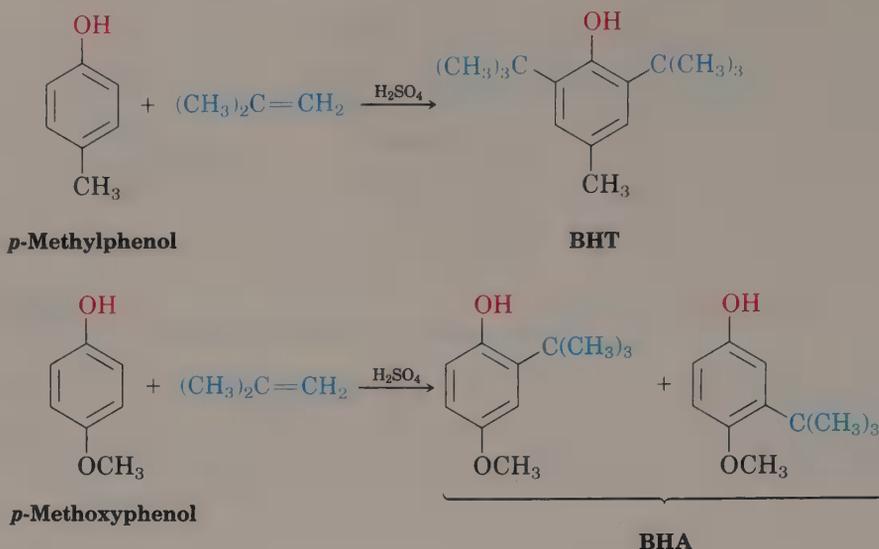


**Figure 26.4** Mechanism of the formation of phenol by acid-catalyzed reaction of cumene hydroperoxide.

In addition to its use in resins and adhesives, phenol also serves as starting material for the synthesis of chlorinated phenols and of the food preservatives BHT (butylated hydroxytoluene) and BHA (butylated hydroxyanisole). Thus, pentachlorophenol, a widely used wood preservative, is prepared by reaction of phenol with excess chlorine. The herbicide 2,4-D (2,4-dichlorophenoxyacetic acid) is prepared from 2,4-dichlorophenol, and the hospital antiseptic agent hexachlorophene is prepared from 2,4,5-trichlorophenol.



The food preservative BHT is prepared by Friedel–Crafts alkylation of *p*-methylphenol (*p*-cresol) with 2-methylpropene in the presence of acid; BHA is prepared similarly by alkylation of *p*-methoxyphenol.



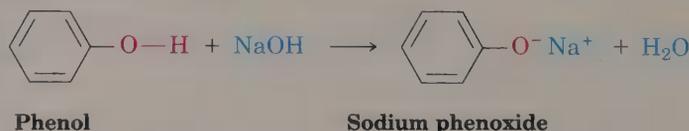
**PROBLEM** .....

- 26.10** 2,4-Dichlorophenoxyacetic acid is prepared from phenol by a two-step sequence involving an electrophilic aromatic substitution followed by a Williamson ether synthesis. Formulate the reactions.
- .....

## 26.6 Properties of Phenols: Acidity

The properties of phenols are similar in many respects to those of alcohols. Low-molecular-weight phenols are generally somewhat water-soluble and are high boiling because of intermolecular hydrogen bonding. The most important property of phenols, however, is their acidity. Phenols are weak acids that dissociate to a small extent in aqueous solution to give  $\text{H}_3\text{O}^+$  and a phenoxide anion,  $\text{ArO}^-$ . Acidity values for some common phenols are given in Table 26.2.

The data in Table 26.2 show that phenols are much more acidic than alcohols. Indeed, some phenols, such as the nitro-substituted ones, even approach or surpass the acidity of carboxylic acids. One practical consequence of this acidity is that phenols are soluble in dilute aqueous sodium hydroxide. Thus, a phenolic component can often be separated from a mixture of compounds simply by basic extraction into aqueous solution, followed by reacidification.



Phenols are more acidic than alcohols because the phenoxide anion is resonance stabilized by the aromatic ring. Sharing of the negative charge

TABLE 26.2 Physical Properties of Some Phenols

Phenol	Melting point (°C)	Boiling point (°C)	pK <sub>a</sub>	
Acetic acid <sup>a</sup>			4.75	
2,4,6-Trinitrophenol	122	—	0.60	
<i>p</i> -Nitrophenol	115	—	7.16	
<i>o</i> -Nitrophenol	97	—	7.21	
<i>m</i> -Nitrophenol	45	216	8.36	
<i>p</i> -Iodophenol	94	—	9.30	
<i>p</i> -Bromophenol	66	238	9.35	
<i>p</i> -Chlorophenol	43	220	9.38	
Phenol	43	182	10.00	
<i>p</i> -Methoxyphenol	57	243	10.21	
<i>p</i> -Methylphenol ( <i>p</i> -cresol)	35	202	10.26	
<i>p</i> -Aminophenol	186	—	10.46	
Ethanol <sup>a</sup>			16.00	

<sup>a</sup>Values for acetic acid and ethanol are given for reference.

on oxygen with the ortho and para positions of the aromatic ring results in increased stability of the phenoxide anion relative to undissociated phenol, and in a consequently lower  $\Delta G^\circ$  for the dissociation reaction.

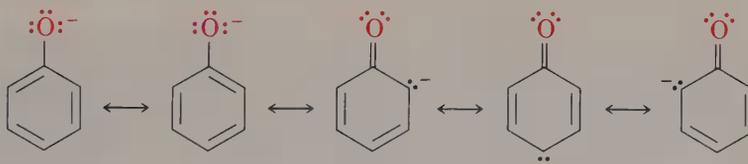
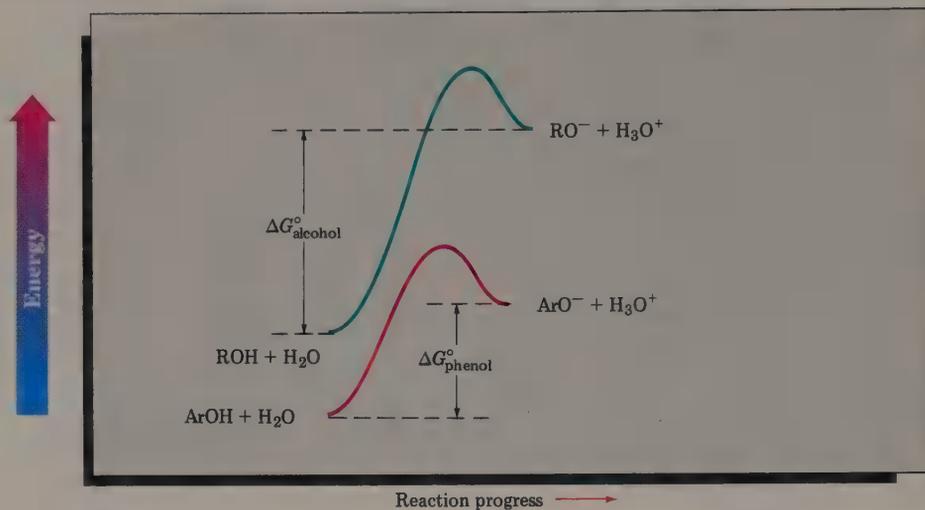


Figure 26.5 compares the acidity of phenols and alcohols on reaction energy diagrams.

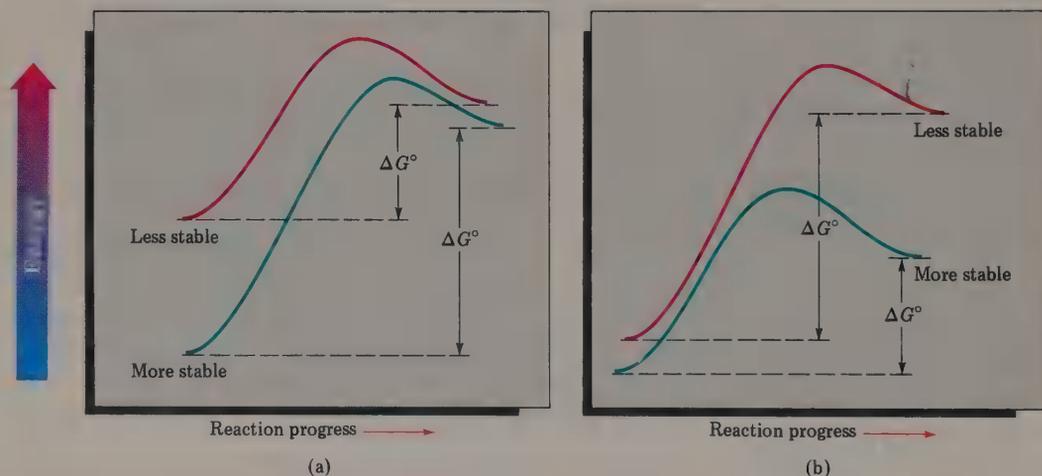
The arguments just used to account for phenol acidity are similar to those used in Section 26.2 to account for arylamine basicity, but the two situations are opposite. Arylamines are less basic than alkylamines because resonance stabilization of the free arylamine is greater than the stabilization of the arylammonium ion. Phenols, however, are more acidic than alcohols because resonance stabilization of the phenoxide ion is greater than that of the free phenol.

In general, any effect that stabilizes the starting material more than it stabilizes the product will raise  $\Delta G^\circ$  and disfavor the reaction. Conversely, any effect that stabilizes the product more than it stabilizes the starting material will lower  $\Delta G^\circ$  and favor the reaction. These effects are shown in Figure 26.6.

Substituted phenols can be either more acidic or less acidic than phenol itself, depending on their structure. As a general rule, phenols with an electron-withdrawing substituent are more acidic, since these substituents stabilize the phenoxide ion by delocalizing the negative charge. Phenols

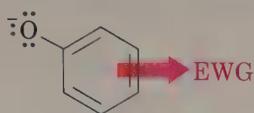


**Figure 26.5** A comparison of phenol and alcohol acidities. Phenols (red curve) are more acidic than alcohols (green curve) because a phenoxide ion is stabilized relative to free phenol more than an alkoxide ion is stabilized relative to free alcohol.

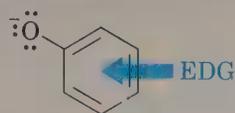


**Figure 26.6** Effect of stabilizing factors on equilibria ( $\Delta G^\circ$ ). (a) Stabilization of starting material relative to product raises  $\Delta G^\circ$  and disfavors the reaction, whereas (b) stabilization of product relative to starting material lowers  $\Delta G^\circ$  and favors the reaction.

with an electron-donating substituent, however, are less acidic, since these substituents destabilize the phenoxide ion by concentrating the charge.

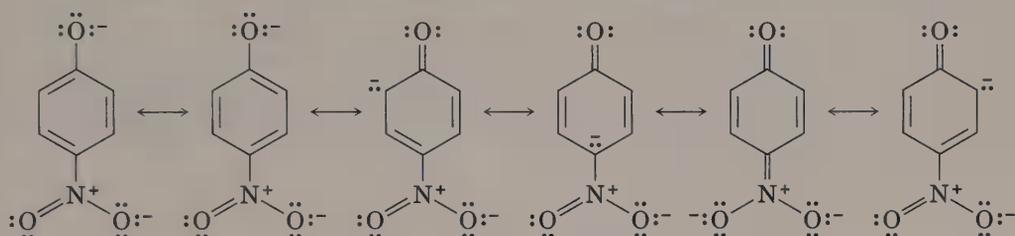


Electron-withdrawing groups (EWG) stabilize phenoxide anion, resulting in increased phenol acidity



Electron-donating groups (EDG) destabilize phenoxide anion, resulting in decreased phenol acidity

The acidifying effect of an electron-withdrawing substituent is particularly noticeable for phenols having a nitro group at the ortho or para positions because the phenoxide anion is strongly stabilized in these cases.



Note that the effect of substituents on phenol acidity is the same as their effect on benzoic acid acidity (Section 20.5), but *opposite* to their effect on arylamine basicity (Table 26.3). The same resonance factors operate in all three cases.

**TABLE 26.3** Substituent Effect on Acidity of Benzoic Acids and Phenols, and on Basicity of Arylamines

Type of compound	Effect of group Y
	Electron-withdrawing; increased acidity Electron-donating; decreased acidity
	Electron-withdrawing; increased acidity Electron-donating; decreased acidity
	Electron-withdrawing; decreased basicity Electron-donating; increased basicity

## PROBLEM .....

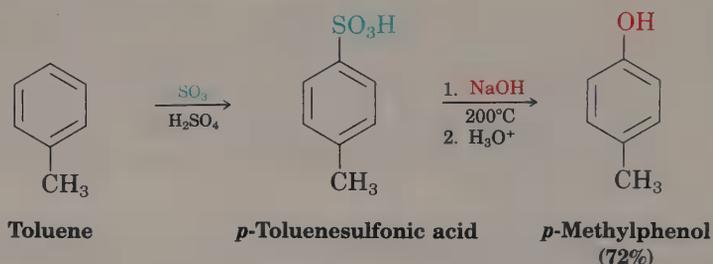
26.11 Rank the compounds in each group in order of increasing acidity:

- (a) Phenol, *p*-methylphenol, *p*-(trifluoromethyl)phenol  
 (b) Benzyl alcohol, phenol, *p*-hydroxybenzoic acid  
 (c) *p*-Bromophenol, 2,4-dibromophenol, 2,4,6-tribromophenol
- .....

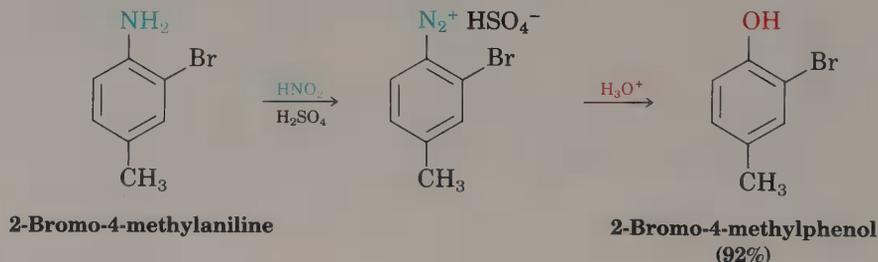
## 26.7 Preparation of Phenols

We've already seen the two best methods of phenol synthesis. To review briefly:

- Alkali fusion of aromatic sulfonates takes place when an arene-sulfonic acid is melted with sodium hydroxide at high temperature (Section 16.2). Few functional groups can survive such harsh conditions, and the reaction is therefore limited to the preparation of alkyl-substituted phenols.



- Hydrolysis of arenediazonium salts in a Sandmeyer-type replacement reaction is the most versatile and widely used laboratory method of phenol synthesis (Section 26.4). Most functional groups are compatible with the reaction conditions needed, and yields are generally good.



## PROBLEM .....

26.12 *p*-Cresol (*p*-methylphenol) is used industrially both as an antiseptic and as a starting material to prepare the food additive BHT. How would you prepare *p*-cresol from benzene?

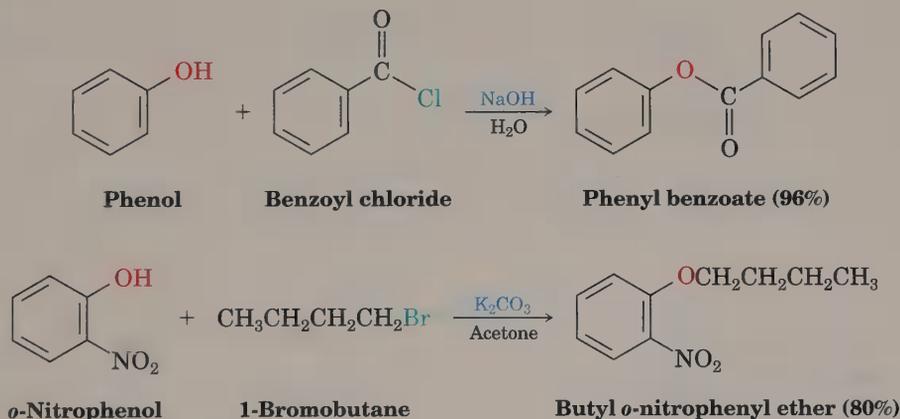
## PROBLEM.....

- 26.13 Carvacrol (5-isopropyl-2-methylphenol) is a natural product isolated from oregano, thyme, and marjoram. Propose two different syntheses of carvacrol from benzene.
- .....

## 26.8 Reactions of Phenols

### ALCOHOL-LIKE REACTIONS

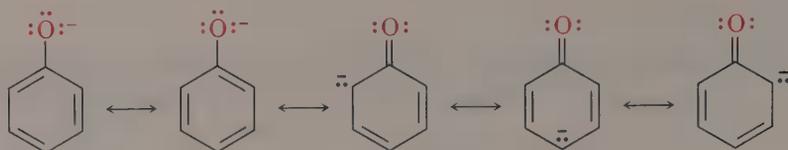
Much of the chemistry of phenols is similar to that of alcohols. Thus, phenols can be converted into esters by reaction with acid chlorides or acid anhydrides, and into ethers by reaction with alkyl halides in the presence of base (Williamson synthesis). Both reactions occur under fairly mild conditions since phenols are so much more acidic than alcohols and since the reactive phenoxide ion intermediates are formed much more readily than alkoxide ions. Direct Fischer esterification by acid-catalyzed reaction between a phenol and a carboxylic acid, however, is not usually successful.



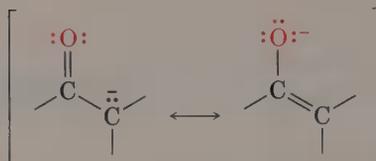
### ELECTROPHILIC AROMATIC SUBSTITUTION REACTIONS

The hydroxyl group is a strongly activating, ortho- and para-directing substituent in electrophilic aromatic substitution reactions (Section 16.5). As a result, phenols are highly reactive substrates for electrophilic halogenation, nitration, and sulfonation, as well as for coupling with diazonium salts to produce azo dyes.

Not surprisingly, phenoxide ions are even more reactive toward electrophilic aromatic substitution than neutral phenols because they have a full negative charge:

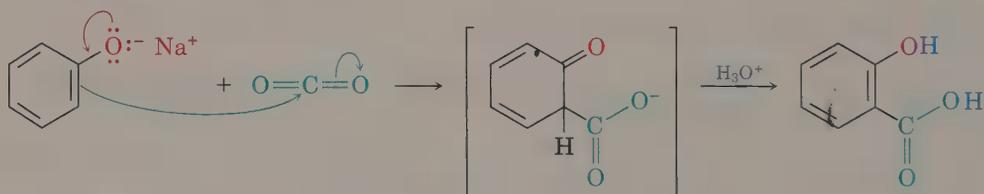


Phenoxide ion



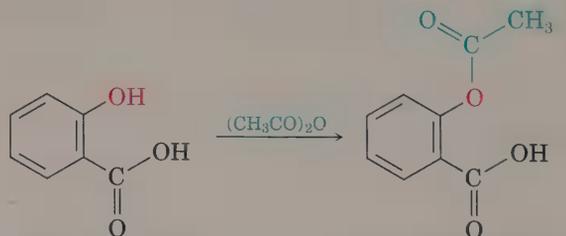
An enolate ion

Resonance structures of a phenoxide anion show a similarity to the resonance structures of a ketone enolate anion, which suggests the possibility that phenoxides might undergo alpha-substitution reactions similar to those of ketones (Section 22.6). In practice, phenoxide ions are less reactive than enolate ions because of the stability of the benzene ring. Nevertheless, there are a number of examples of enolate-like reactivity of phenoxide ions. For example, in the **Kolbe–Schmitt<sup>3,4</sup> carboxylation reaction**, a phenoxide ion adds to carbon dioxide under pressure to yield an intermediate keto acid anion that enolizes to give an *o*-hydroxybenzoic acid (salicylic acid). This reaction is the key step in the industrial synthesis of aspirin (acetylsalicylic acid).



Sodium phenoxide

Salicylic acid



Salicylic acid

Acetylsalicylic acid  
(Aspirin)

<sup>3</sup>Herman Kolbe (1818–1884); b. Germany; Ph.D. Göttingen; professor, universities of Marburg and Leipzig.

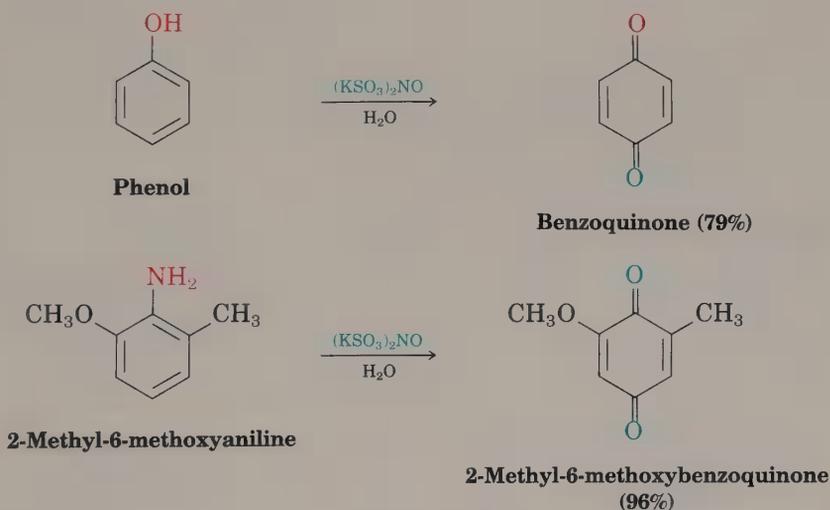
<sup>4</sup>Rudolf Schmitt (1830–1898); b. Wippershain, Germany; Ph.D. Marburg; professor, University of Dresden.

## PROBLEM .....

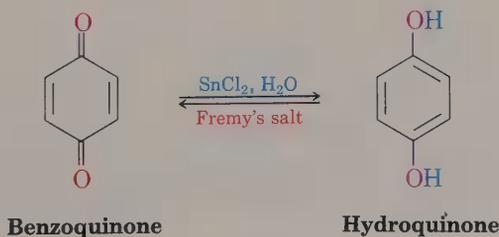
- 26.14 When sodium phenoxide is treated with allyl bromide in a Williamson ether synthesis, a mixture of phenyl allyl ether and *o*-allylphenol is formed. How can you account for the formation of *o*-allylphenol?
- .....

## OXIDATION OF PHENOLS: QUINONES

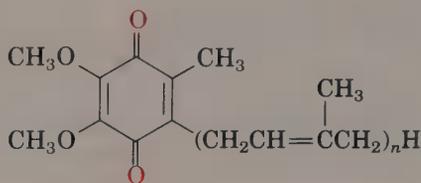
The susceptibility of phenols to electrophilic aromatic substitution is one consequence of the electron-rich nature of the phenol ring. Another consequence is the susceptibility of phenols to oxidation. Treatment of a phenol with any of a number of strong oxidizing agents yields a 2,5-cyclohexadiene-1,4-dione, or **quinone**. Older procedures employed sodium dichromate as oxidant, but Fremy's salt, potassium nitrosodisulfonate  $[(\text{KSO}_3)_2\text{NO}]$ , is now preferred. The reaction takes place under mild conditions through a radical mechanism, and good yields are normally obtained. Arylamines are similarly oxidized to quinones.



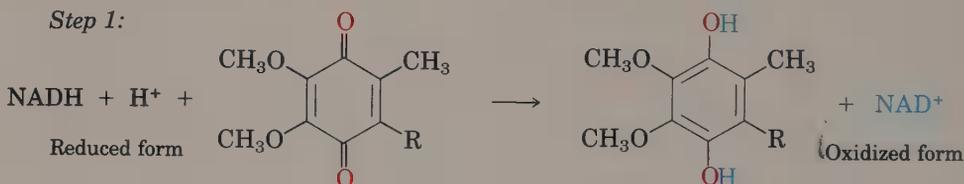
Quinones are an interesting and valuable class of compounds because of their oxidation–reduction properties. They can be easily reduced to **hydroquinones** (*p*-dihydroxybenzenes) by reagents like  $\text{NaBH}_4$  and  $\text{SnCl}_2$ , and hydroquinones can be easily reoxidized back to quinones by Fremy's salt.



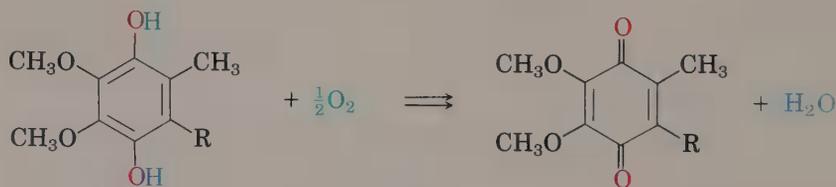
These redox properties of quinones are important to the functioning of living cells, where compounds called *ubiquinones* act as biochemical oxidizing agents to mediate the electron-transfer processes involved in energy production. Ubiquinones, also called *coenzymes Q*, are components of the cells of all aerobic organisms, from the simplest bacterium to humans. They are so named because of their ubiquitous occurrence in nature.

Ubiquinones ( $n = 1-10$ )

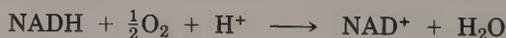
Ubiquinones function within the mitochondria of cells as mobile electron-carriers (oxidizing agents) to mediate the respiration process in which electrons are transported from the biological reducing agent NADH (reduced form of nicotinamide adenine dinucleotide) to molecular oxygen. Although a complex series of steps is involved in the overall process, the ultimate result is a cycle whereby NADH is oxidized to  $\text{NAD}^+$ , oxygen is reduced to water, and energy is produced. Ubiquinone acts only as an intermediary and is itself unchanged.



Step 2:

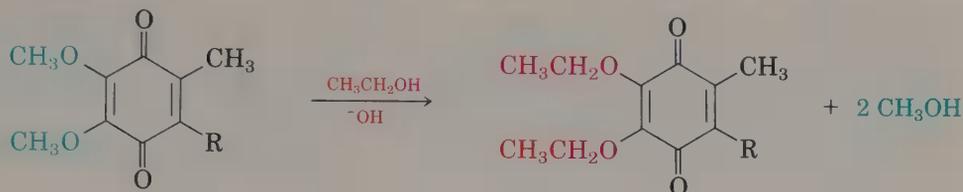


Net change:



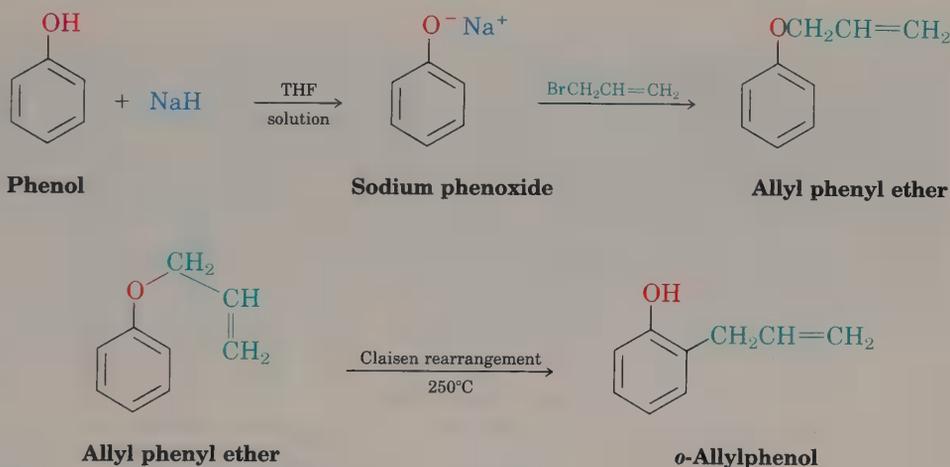
**26.15** Early work on the structural elucidation of ubiquinones was complicated by the fact that extraction of the compounds from cells was carried out using basic ethanol

solution. Under these conditions, the ubiquinone methoxyl groups became exchanged for ethoxyls. Propose a mechanism to account for this exchange. (See Section 19.18.)



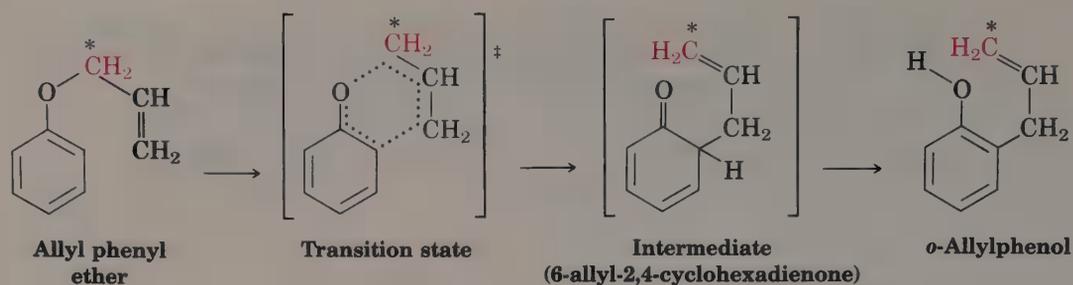
## CLAISEN REARRANGEMENT

Treatment of a phenoxide ion with 3-bromopropene results in a Williamson ether synthesis and production of an allyl phenyl ether. Heating the allyl phenyl ether to 200–250°C then effects rearrangement leading to an *o*-allylphenol (the **Claisen rearrangement**). The net effect is alkylation of the ortho position of the phenol.



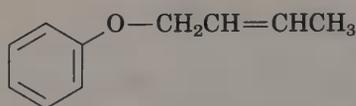
Claisen rearrangement of allyl phenyl ethers to yield *o*-allylphenols is a general reaction that is compatible with the presence of many other substituents on the benzene ring. The reaction proceeds through a pericyclic mechanism in which a concerted reorganization of bonding electrons occurs via a cyclic six-membered-ring transition state. The 2-allylcyclohexadienone intermediate then tautomerizes to *o*-allylphenol.

Good evidence for this mechanism comes from the observation that the rearrangement takes place with an *inversion* of the allyl unit. That is, allyl phenyl ether containing a  $^{14}\text{C}$  label on the allyl ether carbon atom yields *o*-allylphenol in which the label is on the terminal carbon. It would be very difficult to explain this result by any mechanism other than a pericyclic one, which we'll look at in more detail in Section 30.10.



## PROBLEM

- 26.16 What product would you expect to obtain from Claisen rearrangement of 2-butenyl phenyl ether?



2-Butenyl phenyl ether



## 26.9 Spectroscopy of Arylamines and Phenols

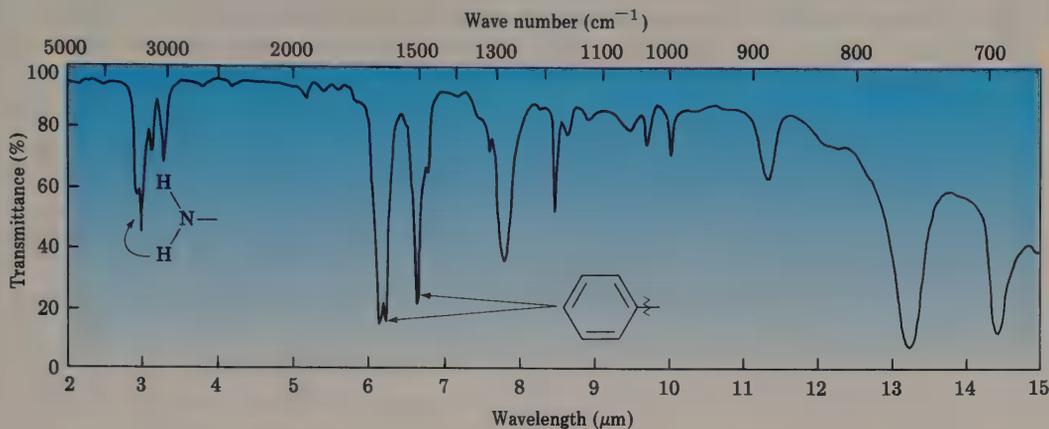
### INFRARED SPECTROSCOPY

The infrared spectra of arylamines and phenols are little different from those of aliphatic amines and alcohols. Thus, aniline shows the usual N–H infrared absorptions at 3400 and 3500  $\text{cm}^{-1}$  characteristic of a primary amine, as well as a pair of bands at 1500 and 1600  $\text{cm}^{-1}$  characteristic of aromatic rings (Figure 26.7). Note that the infrared spectrum of aniline also shows the typical monosubstituted aromatic-ring peaks at 690 and 760  $\text{cm}^{-1}$ .

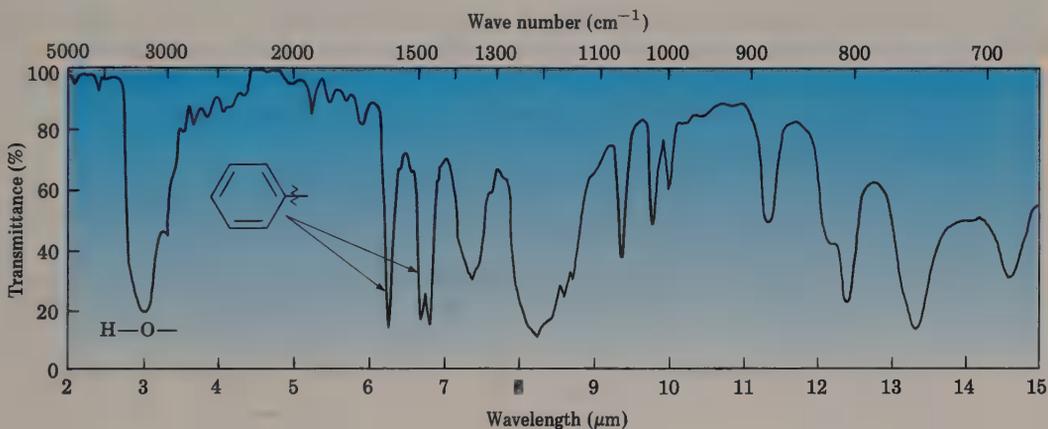
Phenol shows a characteristic broad absorption at 3500  $\text{cm}^{-1}$  due to the hydroxyl group, and the usual 1500 and 1600  $\text{cm}^{-1}$  aromatic bands (Figure 26.8). Here, too, the monosubstituted aromatic-ring peaks at 690 and 760  $\text{cm}^{-1}$  are visible.

### NUCLEAR MAGNETIC RESONANCE SPECTROSCOPY

Arylamines and phenols, like all aromatic compounds, show  $^1\text{H}$  NMR absorptions near 7–8  $\delta$ , the expected position for aromatic-ring protons. In addition, amine N–H protons usually absorb in the 2–3  $\delta$  range, and phenol O–H protons absorb at 3–8  $\delta$ . In neither case are these absorptions uniquely diagnostic for arylamines or phenols, since other kinds of protons absorb in the same range. As a result, a combination of both NMR and infrared evidence is usually needed to assign structure.



**Figure 26.7** Infrared spectrum of aniline.



**Figure 26.8** Infrared spectrum of phenol.

As was true for alcohols (Section 17.11), the identity of the NMR peak due to N-H and O-H protons is easily determined by adding a small amount of  $D_2O$  to the sample tube. Since the O-H and N-H protons rapidly exchange with added  $D_2O$ , their peaks disappear from the spectrum.

## 26.10 Summary and Key Words

**Arylamines**, like their aliphatic counterparts, are basic. The base strength of arylamines is generally lower than that of aliphatic amines, though, because the nitrogen lone-pair electrons are delocalized into the aromatic

ring by orbital overlap with the aromatic pi system. As a general rule, electron-withdrawing substituents on the ring further weaken the basicity of a substituted aniline, while electron-donating substituents increase basicity.

Substituted anilines are almost always prepared by nitration of the appropriate aromatic ring, followed by reduction. An amine group is a strongly activating, ortho- and para-directing substituent, and electrophilic aromatic substitution is an important reaction of arylamines. If the amine group makes the ring too reactive, however, its reactivity can be modulated by converting it into a nonbasic amide.

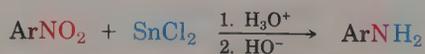
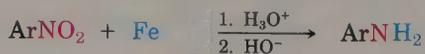
The most important reaction of arylamines is conversion by nitrous acid into **arenediazonium salts**,  $\text{ArN}_2^+ \text{X}^-$ . The diazonio group can then be replaced by many other substituents in the **Sandmeyer reaction** to give a wide variety of substituted aromatic compounds. For example, aryl chlorides, bromides, iodides, and nitriles can be prepared from arenediazonium salts, as can arenes and phenols. In addition to their reactivity toward substitution reactions, diazonium salts undergo coupling with phenols and arylamines to give brightly colored azo dyes.

**Phenols** are aromatic counterparts of alcohols but are much more acidic, since phenoxide anions can be stabilized by delocalization of the negative charge into the aromatic ring. Substitution of the aromatic ring by an electron-withdrawing group increases phenol acidity, and substitution by an electron-donating group decreases acidity. Phenols are generally prepared either by alkali fusion of an aromatic sulfonate or by hydrolysis of an arenediazonium salt.

Reactions of phenols can occur either at the hydroxyl group or on the aromatic ring. For example, a phenol hydroxyl can be converted into an ester or an ether group. Phenyl allyl ethers are particularly interesting since they undergo **Claisen rearrangement** to give *o*-allylphenols when heated to 250°C. The hydroxyl group strongly activates the aromatic ring toward electrophilic substitution reactions. In addition, phenols can be oxidized to **quinones** by reaction with Fremy's salt, potassium nitrosodisulfonate.

## 26.11 Summary of Reactions

1. Preparation of arylamines
  - a. Reduction of nitrobenzenes (Section 26.2)



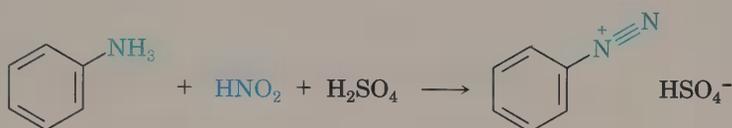
## 2. Reactions of arylamines

## a. Electrophilic aromatic substitution (Sections 16.2 and 26.3)



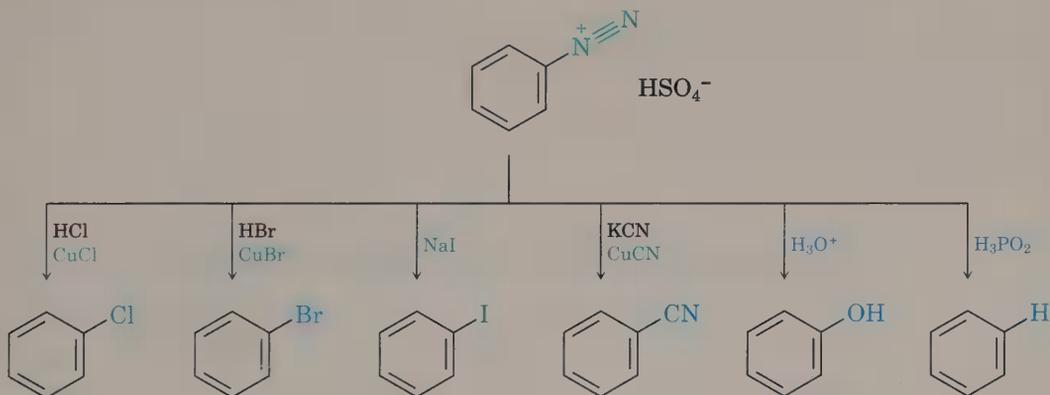
Ortho- and para-directing

## b. Formation of arenediazonium salts (Section 26.3)

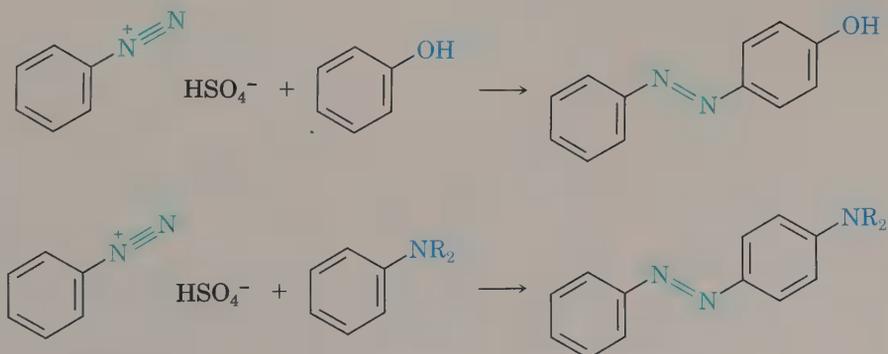


## c. Reaction of arenediazonium salts (Section 26.3)

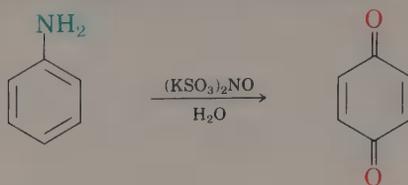
## (1) Sandmeyer-type reactions



## (2) Diazonium coupling



## d. Oxidation to quinones (Section 26.8)

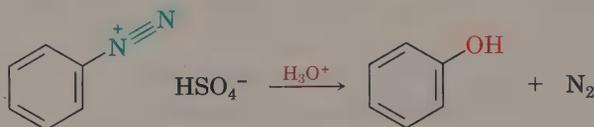


## 3. Preparation of phenols

## a. Alkali fusion of aryl sulfonates (Sections 16.2 and 26.7)

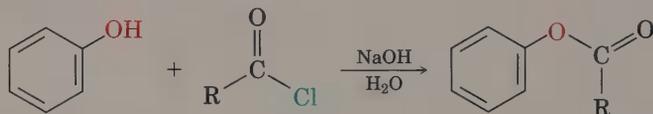


## b. Hydrolysis of arenediazonium salts (Section 26.3)

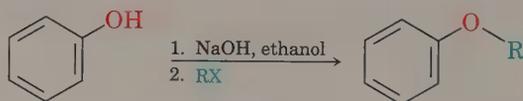


## 4. Reactions of phenols

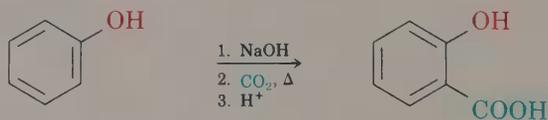
## a. Ester formation (Section 26.8)



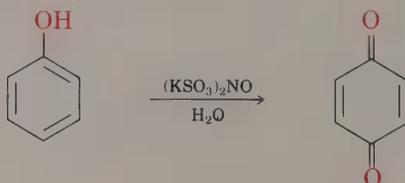
## b. Williamson ether synthesis (Sections 18.4 and 26.8)



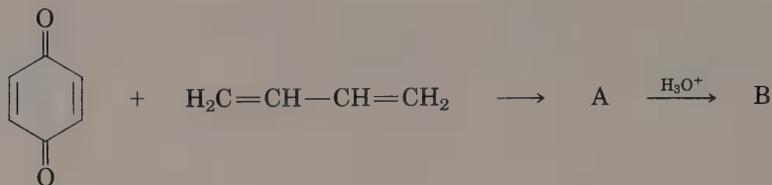
## c. Kolbe-Schmitt carboxylation (Section 26.8)



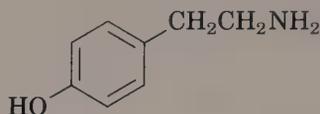
## d. Oxidation to quinones (Section 26.8)





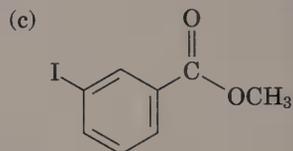
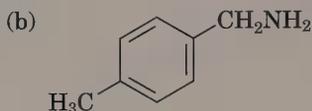
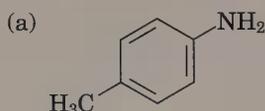


- 26.24** Tyramine is an alkaloid found, among other places, in mistletoe and ripe cheese. How would you synthesize tyramine from benzene? How would you synthesize it from toluene?



Tyramine

- 26.25** How can you account for the fact that diphenylamine does not dissolve in dilute aqueous HCl and appears to be nonbasic?
- 26.26** How would you prepare these compounds from toluene? A diazonio replacement reaction is needed in some instances.

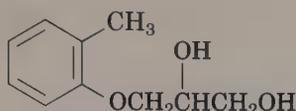


- 26.27** Show the products from reaction of *p*-bromoaniline with these reagents:
- |  |  |
|--|--|
| (a) excess $\text{CH}_3\text{I}$           | (b) $\text{HCl}$   |
| (c) $\text{NaNO}_2, \text{H}_2\text{SO}_4$ | (d) $\text{CH}_3\text{COCl}$   |
| (e) $\text{CH}_3\text{MgBr}$               | (f) $\text{CH}_3\text{CH}_2\text{Cl}, \text{AlCl}_3$                     |
| (g) Product of (c) with $\text{CuCl}$      | (h) Product of (d) with $\text{CH}_3\text{CH}_2\text{Cl}, \text{AlCl}_3$ |
- 26.28** Show the products from reaction of *o*-chlorophenol with these reagents:
- |  |   |
|--|---|
| (a) $\text{NaOH}$ , then $\text{CH}_3\text{I}$ | (b) $\text{CH}_3\text{COCl}$ , pyridine                         |
| (c) Fremy's salt                               | (d) $\text{CH}_3\text{CH}_2\text{CH}_2\text{Cl}, \text{AlCl}_3$ |
- 26.29** Reaction of anthranilic acid (*o*-aminobenzoic acid) with  $\text{NaNO}_2$  and  $\text{H}_2\text{SO}_4$  yields a diazonium salt that can be treated with base to yield a neutral diazonium carboxylate.
- (a) What is the structure of the neutral diazonium carboxylate?
- (b) Heating the diazonium carboxylate results in the formation of  $\text{CO}_2$ ,  $\text{N}_2$ , and a high-energy intermediate that reacts with 1,3-cyclopentadiene to yield the following organic product:



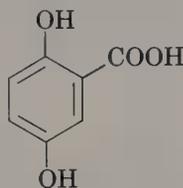
What is the structure of the reactive intermediate, and what kind of reaction does it undergo with cyclopentadiene?

- 26.30** Mephesisin is a drug used as a muscle relaxant and sedative. Propose a synthesis of mephesisin from benzene and any other reagents needed.



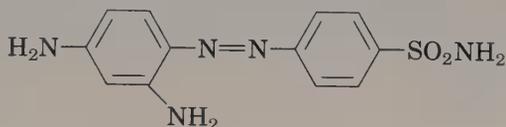
Mephesisin

- 26.31 How would you prepare these substances?  
 (a) *p*-Iodobenzoic acid from aniline  
 (b) *o*-Iodobromobenzene from benzene
- 26.32 Gentisic acid is a naturally occurring hydroquinone found in gentian. Its sodium salt is used medicinally as an antirheumatic agent. How would you prepare gentisic acid from benzene?



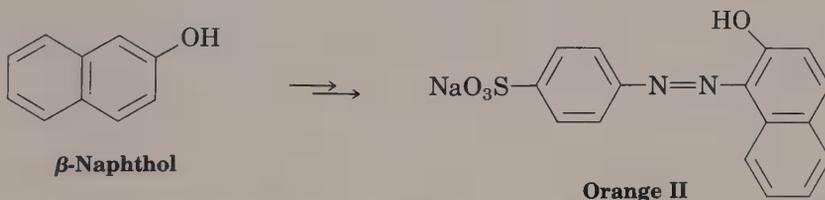
Gentisic acid

- 26.33 Prontosil is an antibacterial azo dye that was once used for urinary tract infections. How would you prepare prontosil from benzene?



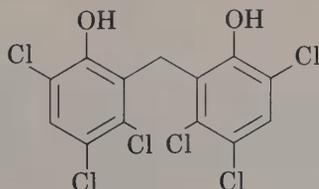
Prontosil

- 26.34 How would you synthesize the dye Orange II from benzene and  $\beta$ -naphthol?

 $\beta$ -Naphthol

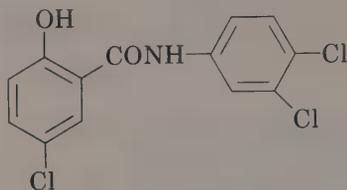
Orange II

- 26.35 2-Nitro-3,4,6-trichlorophenol is used as a lampricide—a compound toxic to lampreys—to combat the intrusion of sea lampreys into the Great Lakes. How would you synthesize this material from benzene?
- 26.36 The germicidal agent hexachlorophene is prepared by condensation of two molecules of 2,4,5-trichlorophenol with one molecule of formaldehyde in the presence of sulfuric acid. Propose a mechanism for this reaction.



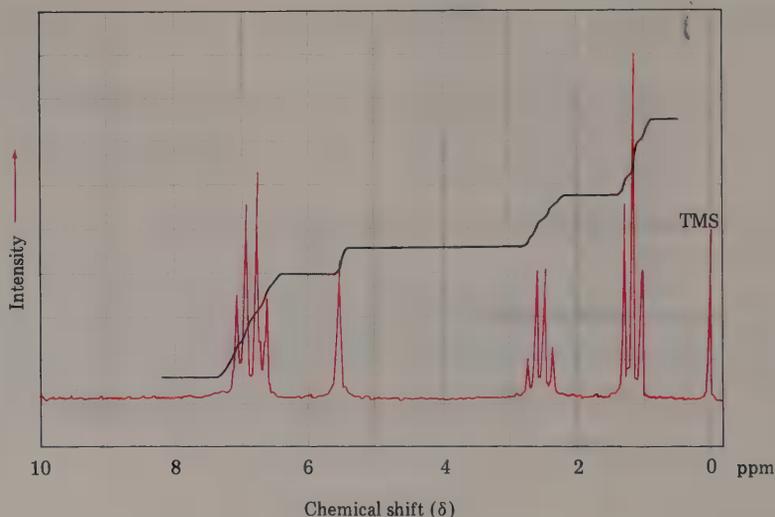
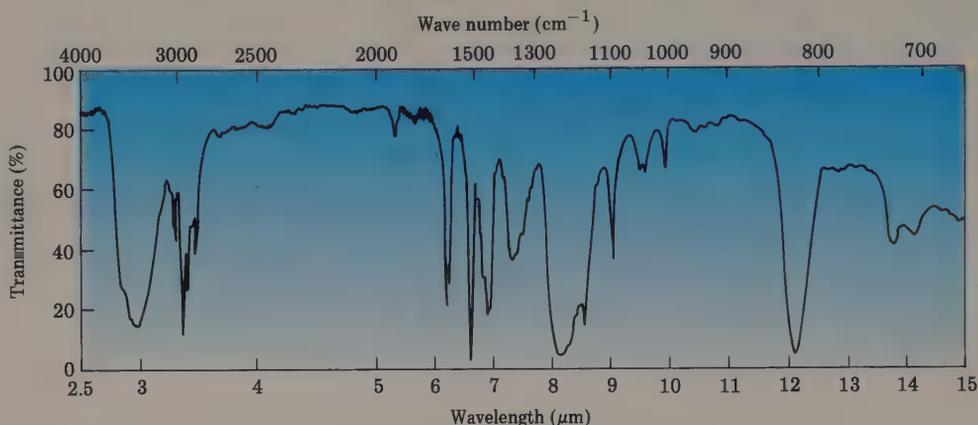
Hexachlorophene

- 26.37 Propose a route from benzene for the synthesis of the antiseptic agent trichlorosalicylanilide.



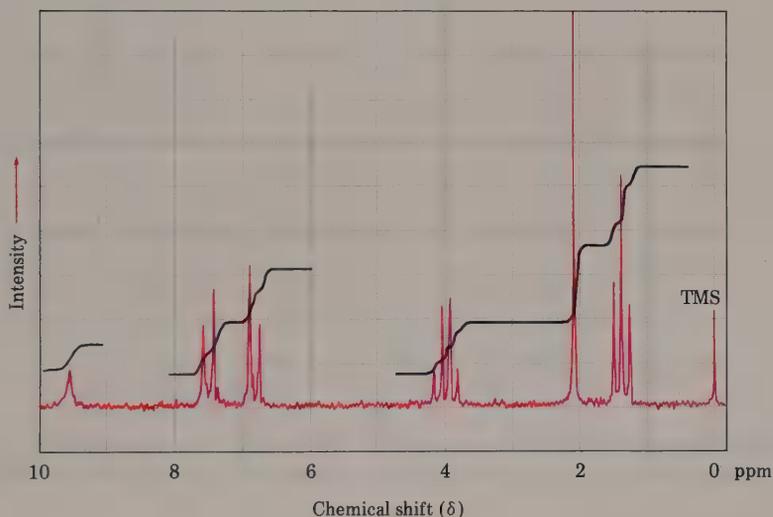
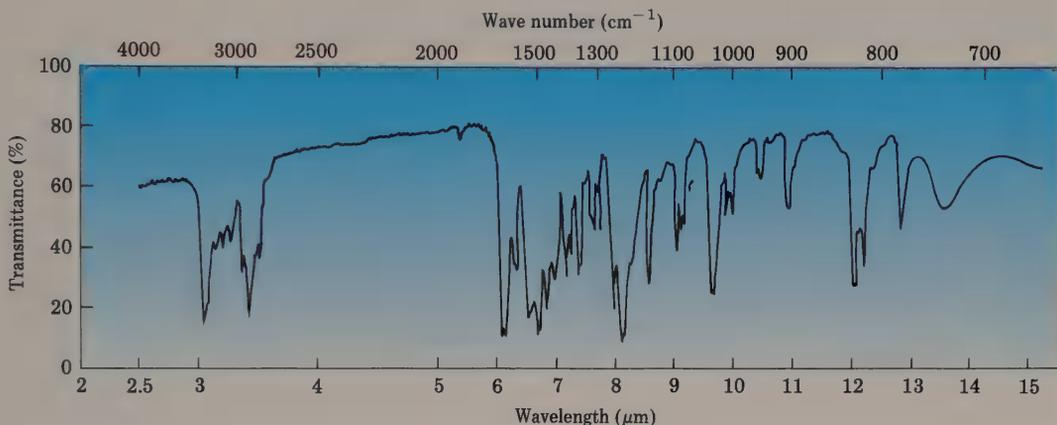
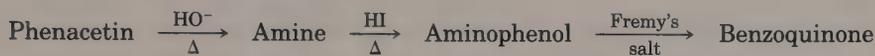
Trichlorosalicylanilide

- 26.38 Compound A,  $C_8H_{10}O$ , has the infrared and  $^1H$  NMR spectra shown. Propose a structure consistent with the observed spectral properties, and assign each peak in the NMR spectrum. Note that the absorption at  $6.0 \delta$  disappears when  $D_2O$  is added.

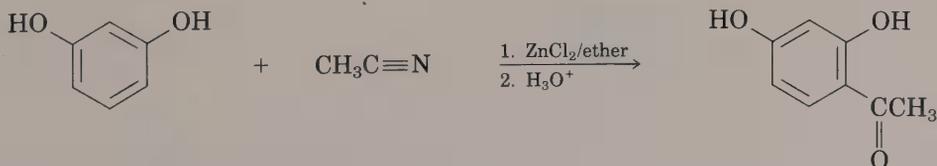


- 26.39 Phenacetin, a substance formerly used in over-the-counter pain relievers, has the formula  $C_{10}H_{13}NO_2$  and the infrared and NMR spectra shown. Phenacetin itself is neutral and does not dissolve in either acid or base. When warmed with aqueous hydroxide, phenacetin yields an amine,  $C_8H_{11}NO$ . When heated with HI, the amine is cleaved to an aminophenol  $C_6H_7NO$ , which, on treatment with Fremy's salt, yields

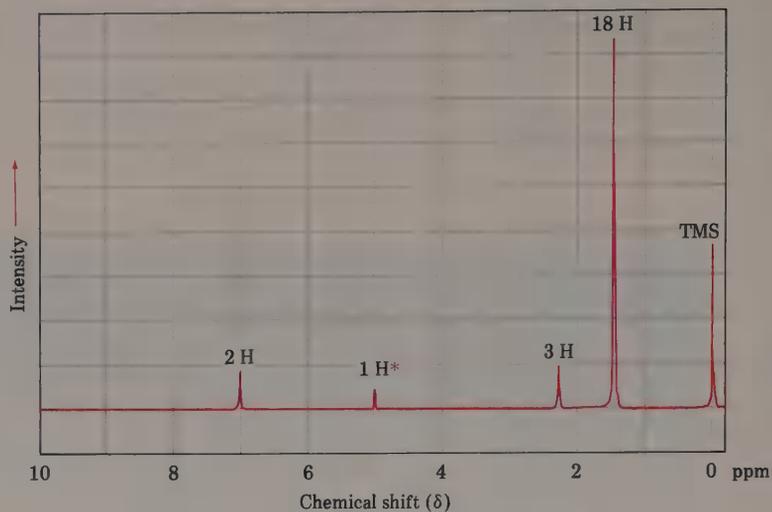
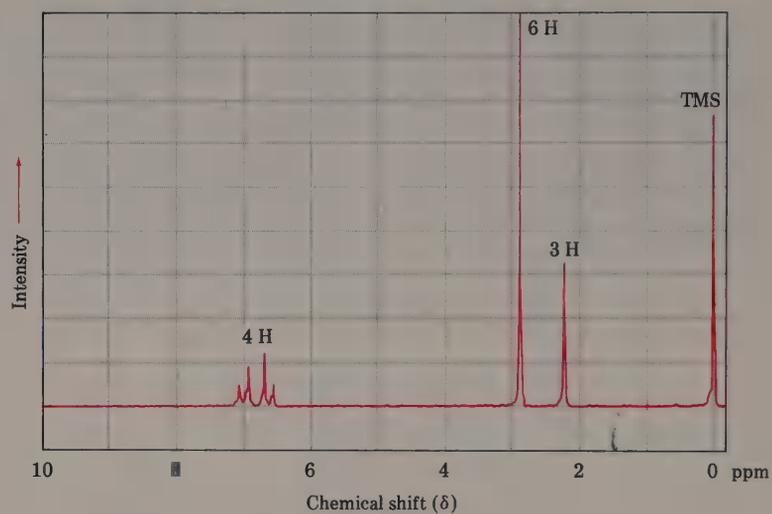
benzoquinone. What is the structure of phenacetin, and what are the structures of the amine and the aminophenol?



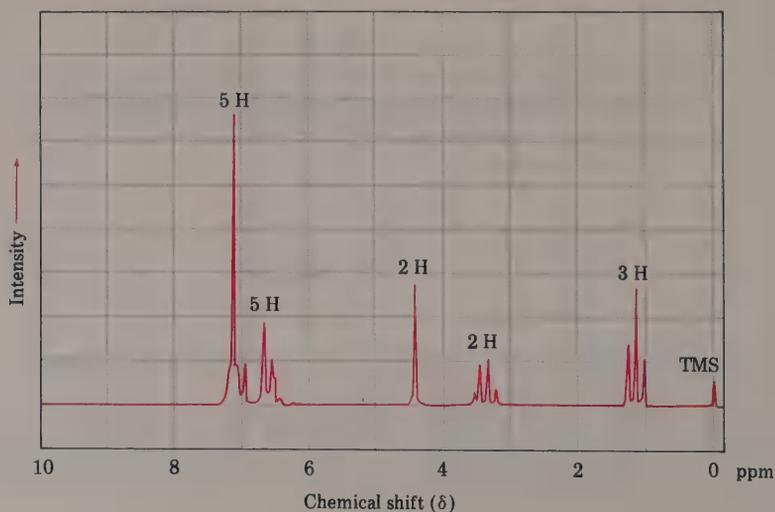
- 26.40** In the Hoesch reaction, resorcinol (*m*-dihydroxybenzene) is treated with a nitrile in the presence of a Lewis acid catalyst. After hydrolysis, an acyl resorcinol is isolated. Propose a mechanism for the Hoesch reaction. To what other well-known reaction is this similar?

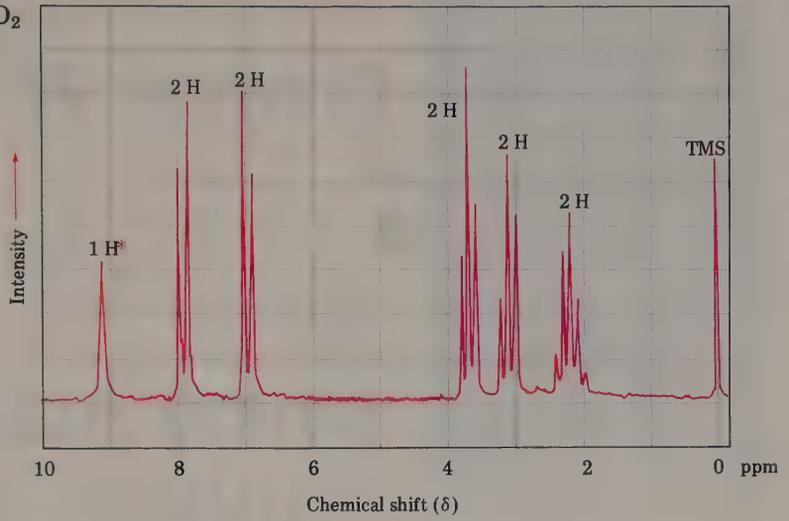


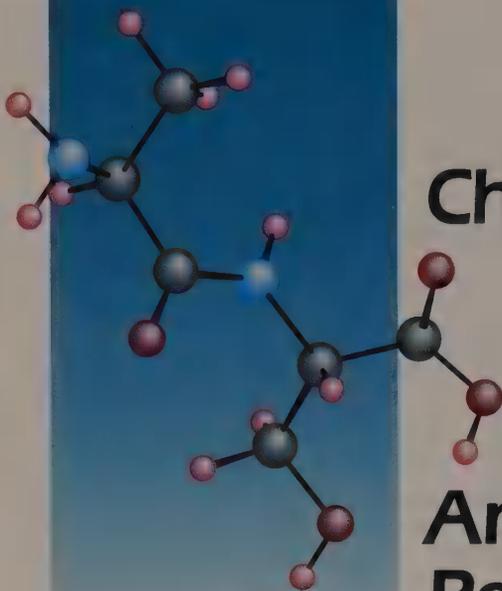
- 26.41** Propose structures for compounds that show the following  $^1\text{H}$  NMR spectra. The peak marked by an asterisk disappears when  $\text{D}_2\text{O}$  is added to the sample.

(a)  $C_{15}H_{24}O$ (b)  $C_9H_{13}N$ 

**26.42** Propose structures for compounds that show the following  $^1H$  NMR spectra. The peak marked by an asterisk disappears when  $D_2O$  is added to the sample.

(a)  $C_{15}H_{17}N$ 

(b)  $C_{10}H_{11}ClO_2$ 

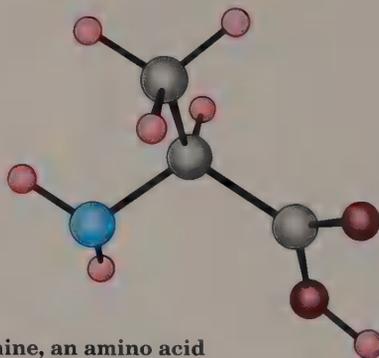
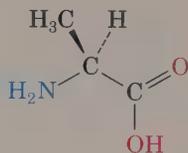


## Chapter 27

# Amino Acids, Peptides, and Proteins

**Proteins** are large biomolecules that occur in every living organism. They're of many different types, and they have many different biological functions. The keratin of skin and fingernails, the fibroin of silk and spider webs, and the collagen of tendons and cartilage are all structural proteins; the insulin that regulates glucose metabolism in the body is a hormonal protein; and the DNA polymerase and reverse transcriptase that serve as biological catalysts to carry out chemical reactions in the cell are proteins called enzymes.

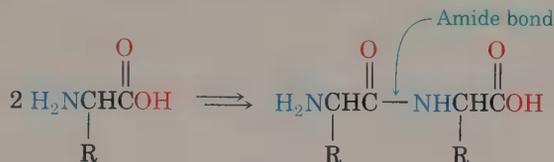
Regardless of their appearance or function, all proteins are chemically similar. All are made up of many **amino acid** units linked together by amide bonds in a long chain. As their name implies, amino acids are difunctional: They contain both a basic amino group and an acidic carboxyl group.



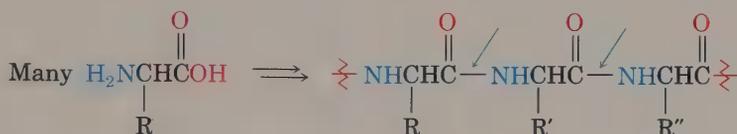
Alanine, an amino acid

The value of amino acids as biological building blocks stems from the fact that they can link together by forming amide, or **peptide**, bonds. A dipeptide results when an amide bond is formed between the  $-\text{NH}_2$  of one amino acid and the  $-\text{COOH}$  of a second amino acid; a tripeptide results from linkage of three amino acids via two amide bonds, and so on. Any number

of amino acids can link together to form large chains. For classification purposes, chains with fewer than 50 amino acids are called **polypeptides**, whereas the term **protein** is reserved for larger chains.



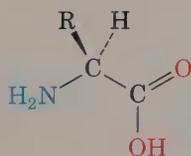
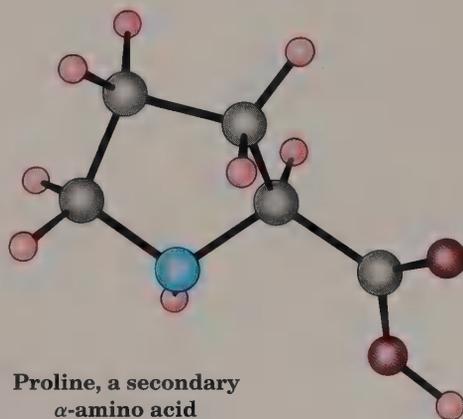
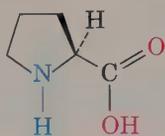
A dipeptide (one amide bond)



A polypeptide (many amide bonds)

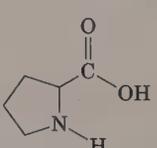
## 27.1 Structures of Amino Acids

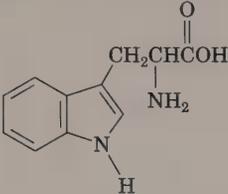
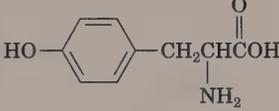
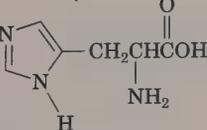
The structures of the 20 amino acids commonly found in proteins are shown in Table 27.1. All 20 are  **$\alpha$ -amino acids**; that is, the amino group in each is a substituent on the carbon atom alpha to (next to) the carbonyl group. Note that 19 of the 20 are primary amines,  $\text{RNH}_2$ , that differ only in the nature of the side-chain substituents but that proline is a secondary amine whose nitrogen and alpha carbon atoms are part of a pyrrolidine ring. Proline can still form amide bonds like the other 19  $\alpha$ -amino acids, however.

A primary  $\alpha$ -amino acid  
(R = ■ side chain)Proline, a secondary  
 $\alpha$ -amino acid

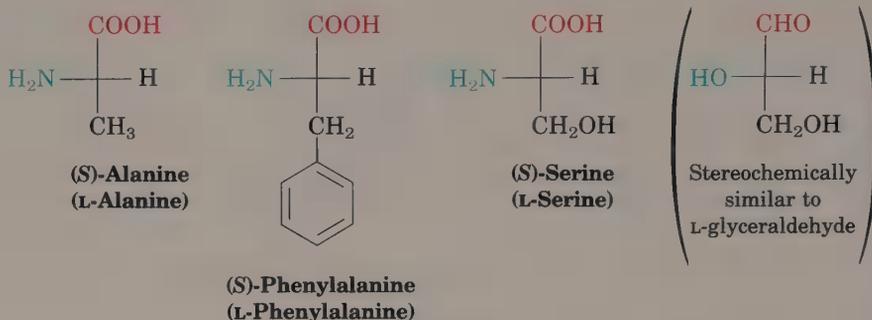
Note also that each of the amino acids in Table 27.1 is referred to by a mnemonic three-letter shorthand code: Ala for alanine, Gly for glycine, and so on. In addition, a one-letter code is used, as shown in parentheses in the table.

**TABLE 27.1** The Twenty Common Amino Acids Found in Proteins; Essential Amino Acids Are Shown in Red

Name	Abbreviations	Molecular weight	Structure	Isoelectric point	pK <sub>a</sub> α-COOH	pK <sub>a</sub> α-NH <sub>3</sub> <sup>+</sup>
<b>Neutral amino acids</b>						
Alanine	Ala (A)	89	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	6.0	2.3	9.8
Asparagine	Asn (N)	132	$\begin{array}{c} \text{O} \quad \quad \text{O} \\ \parallel \quad \quad \parallel \\ \text{H}_2\text{NCCH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.4	2.0	8.8
Cysteine	Cys (C)	121	$\begin{array}{c} \text{O} \\ \parallel \\ \text{HSCH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.0	1.8	10.2
Glutamine	Gln (Q)	146	$\begin{array}{c} \text{O} \quad \quad \quad \text{O} \\ \parallel \quad \quad \quad \parallel \\ \text{H}_2\text{NCCH}_2\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.7	2.2	9.1
Glycine	Gly (G)	75	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_2\text{COH} \\   \\ \text{NH}_2 \end{array}$	6.0	2.3	9.8
Isoleucine	Ile (I)	131	$\begin{array}{c} \text{CH}_3 \quad \quad \text{O} \\   \quad \quad \parallel \\ \text{CH}_3\text{CH}_2\text{CHCHCOH} \\   \\ \text{NH}_2 \end{array}$	6.0	2.3	9.7
Leucine	Leu (L)	131	$\begin{array}{c} \text{CH}_3 \quad \quad \text{O} \\   \quad \quad \parallel \\ \text{CH}_3\text{CHCH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	6.0	2.3	9.7
Methionine	Met (M)	149	$\begin{array}{c} \text{O} \\ \parallel \\ \text{CH}_3\text{SCH}_2\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.7	2.3	9.2
Phenylalanine	Phe (F)	165	$\begin{array}{c} \text{O} \\ \parallel \\ \text{C}_6\text{H}_5\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.9	2.6	9.2
Proline	Pro (P)	115		6.3	2.0	10.6
Serine	Ser (S)	105	$\begin{array}{c} \text{O} \\ \parallel \\ \text{HOCH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	5.7	2.2	9.2

Name	Abbreviations	Molecular weight	Structure	Isoelectric point	pK <sub>a</sub> α-COOH	pK <sub>a</sub> α-NH <sub>3</sub> <sup>+</sup>
Threonine	Thr (T)	119	$\begin{array}{c} \text{OH} \quad \text{O} \\   \quad    \\ \text{CH}_3\text{CHCHCOH} \\   \\ \text{NH}_2 \end{array}$	5.6	2.1	9.1
Tryptophan	Trp (W)	204		5.9	2.4	9.4
Tyrosine	Tyr (Y)	181		5.7	2.2	9.1
Valine	Val (V)	117	$\begin{array}{c} \text{CH}_3 \quad \text{O} \\   \quad    \\ \text{CH}_3\text{CHCHCOH} \\   \\ \text{NH}_2 \end{array}$	6.0	2.3	9.7
<b>Acidic amino acids</b>						
Aspartic acid	Asp (D)	133	$\begin{array}{c} \text{O} \quad \text{O} \\    \quad    \\ \text{HOOCCH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	3.0	2.1	9.8
Glutamic acid	Glu (E)	147	$\begin{array}{c} \text{O} \quad \text{O} \\    \quad    \\ \text{HOOCCH}_2\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	3.2	2.1	9.5
<b>Basic amino acids</b>						
Arginine	Arg (R)	174	$\begin{array}{c} \text{NH} \quad \text{O} \\    \quad    \\ \text{H}_2\text{NCNHCH}_2\text{CH}_2\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	10.8	2.0	9.0
Histidine	His (H)	155		7.6	1.8	9.1
Lysine	Lys (K)	146	$\begin{array}{c} \text{O} \\    \\ \text{H}_2\text{NCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CHCOH} \\   \\ \text{NH}_2 \end{array}$	9.7	2.2	8.9

With the exception of glycine,  $\text{H}_2\text{NCH}_2\text{COOH}$ , the alpha carbons of amino acids are chiral. Two different enantiomeric forms of each amino acid are therefore possible, but nature uses only a single enantiomer to construct proteins. In Fischer projections, naturally occurring amino acids are represented by placing the carboxyl group at the top as if drawing a carbohydrate (Section 24.2) and then placing the amino group on the left. Because of their stereochemical similarity to L sugars—the  $-\text{NH}_2$  group is toward the left (Section 24.3)—the naturally occurring  $\alpha$ -amino acids are often referred to as L-amino acids.



The 20 common amino acids can be further classified as either neutral, acidic, or basic, depending on the nature of their side chains. Fifteen of the twenty have neutral side chains, but two (aspartic acid and glutamic acid) have an extra carboxylic acid function in their side chains, and three (lysine, arginine, and histidine) have basic amino groups in their side chains.

All 20 of the amino acids are necessary for protein synthesis, but humans are thought to be able to synthesize only 10 of the 20. The remaining 10 are called **essential amino acids** because they must be obtained from dietary sources. Failure to include an adequate dietary supply of these essential amino acids can lead to severe deficiency diseases.

PROBLEM .....

- 27.1 Look carefully at the  $\alpha$ -amino acids shown in Table 27.1. How many contain aromatic rings? How many contain sulfur? How many contain alcohols? How many contain hydrocarbon side chains?

PROBLEM .....

- 27.2 Eighteen of the 19 L-amino acids have the *S* configuration at the alpha carbon. Cysteine is the only L-amino acid that has an *R* configuration. Explain.

PROBLEM .....

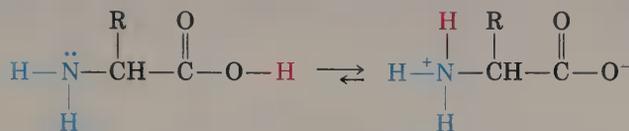
- 27.3 The amino acid threonine, (2*S*,3*R*)-2-amino-3-hydroxybutanoic acid, has two stereogenic centers. Draw a Fischer projection of threonine.

PROBLEM .....

- 27.4 Draw the Fischer projection of a threonine diastereomer, and label its stereogenic centers as *R* or *S* (Problem 27.3).
- .....

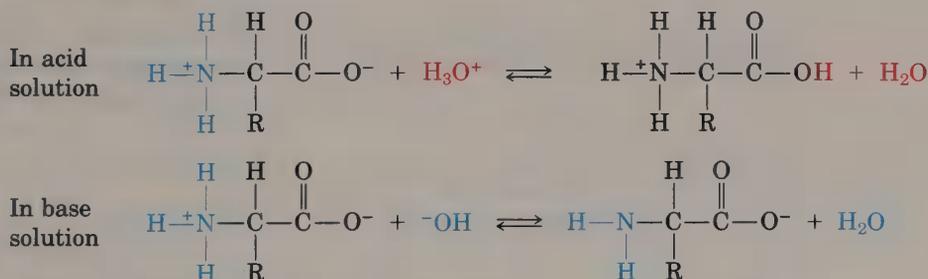
## 27.2 Dipolar Structure of Amino Acids

Since amino acids contain both acidic and basic groups in the same molecule, they undergo an intramolecular acid–base reaction and exist primarily in the form of a dipolar ion, or **zwitterion** (German *zwitter*, “hybrid”):



A zwitterion

Amino acid zwitterions are a kind of internal salt and therefore have many of the physical properties associated with salts. Thus, amino acids have large dipole moments; they're soluble in water but insoluble in hydrocarbons; and they're crystalline substances with high melting points. In addition, amino acids are *amphoteric*: They can react either as acids or as bases, depending on the circumstances. In aqueous acid solution, an amino acid zwitterion can *accept* a proton to yield a cation; in aqueous basic solution, the zwitterion can *lose* a proton to form an anion.



Note that it's the carboxylate anion,  $-\text{COO}^-$ , rather than the amino group that acts as the basic site and accepts the proton in acid solution. Similarly, it's the ammonium cation,  $\text{RNH}_3^+$ , rather than the carboxyl group that acts as the acidic site and donates a proton in base solution.

### PROBLEM .....

**27.5** Draw these amino acids in their zwitterionic forms:

(a) Phenylalanine

(b) Serine

(c) Proline

### PROBLEM .....

**27.6** Write structural formulas for these equations.

(a) Phenylalanine + 1 equiv NaOH  $\longrightarrow$  ?

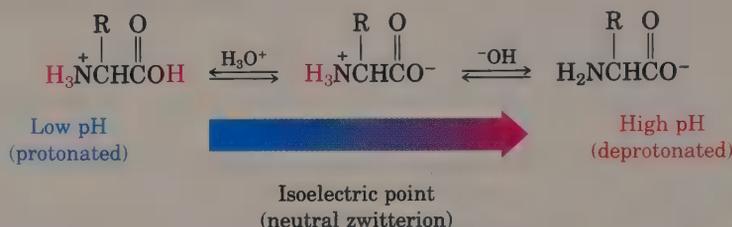
(b) Product of (a) + 1 equiv HCl  $\longrightarrow$  ?

(c) Product of (a) + 2 equiv HCl  $\longrightarrow$  ?

### .....

## 27.3 Isoelectric Point

In acid solution at low pH, an amino acid is protonated and exists primarily as a cation. In basic solution at high pH, an amino acid is deprotonated and exists primarily as an anion. Thus, at some intermediate point, the amino acid must be exactly balanced between anionic and cationic forms and exist primarily as the neutral, dipolar zwitterion. This pH is called the amino acid's **isoelectric point**.

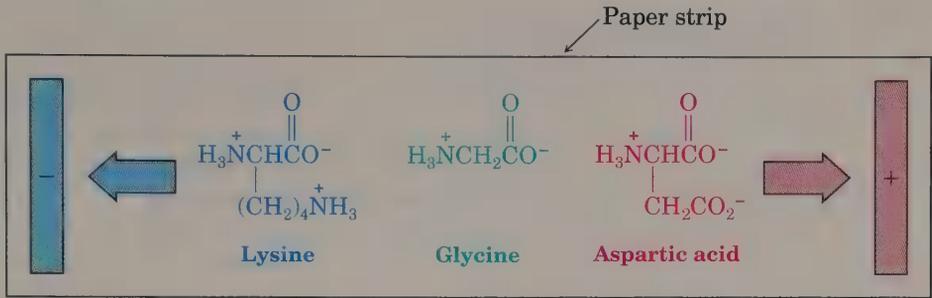


The isoelectric point of an amino acid depends on its structure, with values for the 20 most common amino acids given in Table 27.1. The 15 amino acids with neutral side chains have isoelectric points near neutrality, in the pH range 5.0–6.5. (These values aren't exactly at neutral pH 7, because carboxyl groups are stronger acids in aqueous solution than amino groups are bases.) The two amino acids with acidic side chains have isoelectric points at lower (more acidic) pH, which suppresses dissociation of the extra  $-\text{COOH}$  in the side chain, and the three amino acids with basic side chains have isoelectric points at higher (more basic) pH, which suppresses protonation of the extra amino group. For example, aspartic acid has its isoelectric point at pH 3.0, and lysine has its isoelectric point at pH 9.7.

We can take advantage of the differences in isoelectric points to separate a mixture of amino acids (or a mixture of proteins) into its pure constituents. Using a technique known as **electrophoresis**, a solution of different amino acids is placed near the center of a strip of paper or gel. The paper or gel is moistened with an aqueous buffer of a particular pH, and electrodes are connected to the ends of the strip. When an electric field is applied, those amino acids with negative charges (those that are partially deprotonated because their isoelectric points are below the pH of the buffer) migrate slowly toward the positive electrode. Similarly, those amino acids with positive charges (those that are partially protonated because their isoelectric points are above the pH of the buffer) migrate toward the negative electrode.

Different amino acids migrate at different rates, depending both on their isoelectric point and on the pH of the aqueous buffer. Thus, the different amino acids can be separated. Figure 27.1 illustrates this separation for a mixture of lysine (basic), glycine (neutral), and aspartic acid (acidic).

If specific  $\text{p}K_a$  values are known for the acidic sites of an amino acid (Table 27.1), the exact percentages of protonated, neutral, and deprotonated forms at any pH can be calculated using the **Henderson–Hasselbalch** equation, a simple expression derived by rearranging the standard equation for  $\text{p}K_a$ :

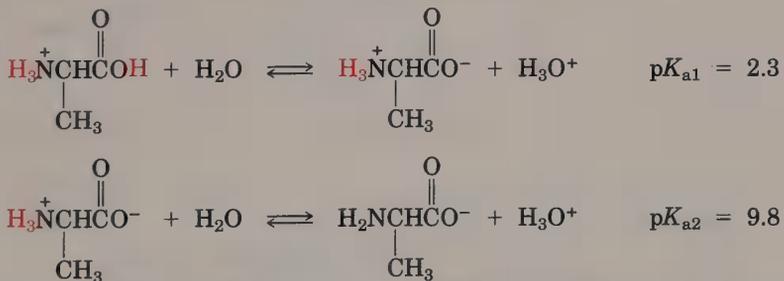


**Figure 27.1** Separation of an amino acid mixture by electrophoresis. At pH = 6.0, glycine molecules are mostly neutral and don't migrate; lysine molecules are largely protonated and migrate toward the negative electrode; and aspartic acid molecules are largely deprotonated and migrate toward the positive electrode. (Lysine has its isoelectric point at 9.7, glycine at 6.0, and aspartic acid at 3.0.)

$$\text{Since } pK_a = -\log \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]} = -\log[\text{H}_3\text{O}^+] - \log \frac{[\text{A}^-]}{[\text{HA}]} = \text{pH} - \log \frac{[\text{A}^-]}{[\text{HA}]}$$

$$\left. \begin{array}{l} \text{then } \text{pH} = pK_a + \log \frac{[\text{A}^-]}{[\text{HA}]} \\ \text{or } \log \frac{[\text{A}^-]}{[\text{HA}]} = \text{pH} - pK_a \end{array} \right\} \text{Henderson-Hasselbalch equation}$$

In other words, the logarithm of the conjugate base concentration divided by the weak acid concentration is equal to the pH minus the  $pK_a$  of the acid. The protonated form of alanine, for example, is a weak acid with  $pK_{a1} = 2.3$ , and the neutral form of alanine is a still weaker acid with  $pK_{a2} = 9.8$ :



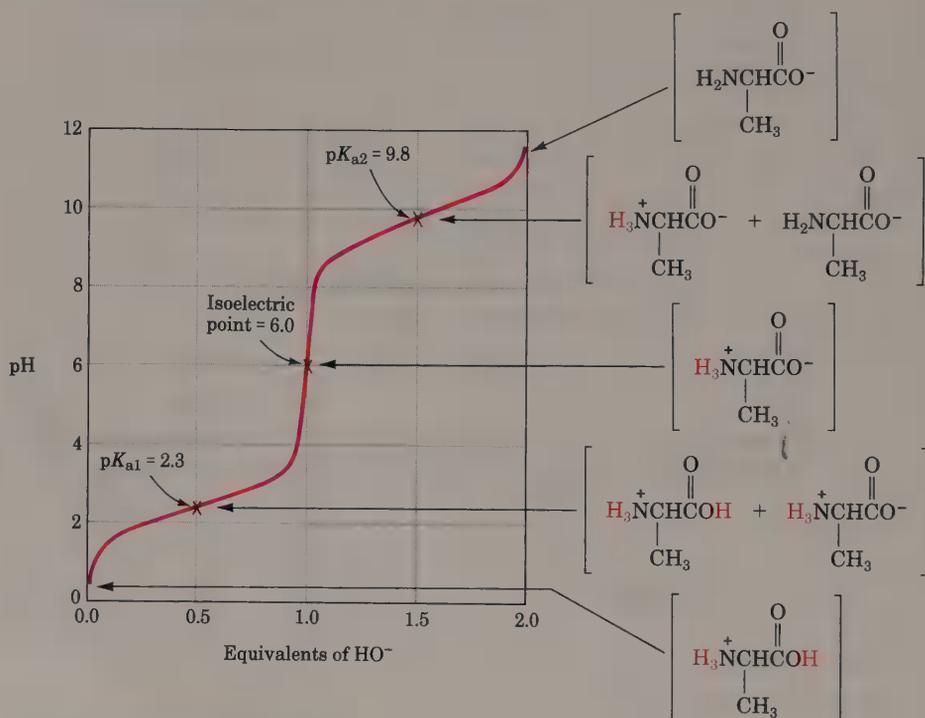
To see how to use the Henderson-Hasselbalch equation, let's find what species are present at pH 1.00:

$$\log \frac{[\text{A}^-]}{[\text{HA}]} = \text{pH} - pK_a = 1.00 - 2.3 = -1.3$$

$$\text{so } \frac{[\text{A}^-]}{[\text{HA}]} = 0.050$$

$$\text{Thus, } [\text{A}^-] = 4.8\% \text{ and } [\text{HA}] = 95.2\%$$

At pH 1, 4.8% of alanine molecules are neutral (zwitterionic) and 95.2% are protonated. Similar calculations can be done at any other pH, leading to the *titration curve* shown in Figure 27.2. Each “leg” of the curve—the first leg from pH 1 to 6 and the second leg from pH 6 to 11—is calculated separately, corresponding to the two dissociations of protonated alanine. Exactly halfway between the two legs is the isoelectric point at 6.0. In essence, it’s as if we started with protonated alanine at low pH and then titrated with NaOH until full deprotonation occurred. When 0.5 equivalent of NaOH is added, the first deprotonation is half complete; when 1.0 equivalent of NaOH is added, the first deprotonation is fully complete and the isoelectric point is reached; when 1.5 equivalents of NaOH are added, the second deprotonation is half complete; and when 2.0 equivalents of NaOH are added, the second deprotonation is complete.



**Figure 27.2** A titration curve for alanine plotted using the Henderson–Hasselbalch equation. Each of the two legs is plotted separately. At pH < 1, alanine is entirely protonated; at pH = 2.3, alanine is a 50:50 mix of protonated and neutral forms; at pH = 6.0, alanine is entirely neutral; at pH = 9.8, alanine is a 50:50 mix of neutral and deprotonated forms; at pH > 11, alanine is entirely deprotonated.

**PROBLEM** .....

- 27.7** Glycine, like alanine, has an isoelectric point at 6.0. Draw structures of the predominant forms of glycine at pH = 2.0, 6.0, and 10.0.

## PROBLEM .....

- 27.8 For the mixtures of amino acids indicated, predict the direction of migration of each component (toward anode or cathode) and the relative rate of migration.
- Valine, glutamic acid, and histidine at pH = 7.6
  - Glycine, phenylalanine, and serine at pH = 5.7
  - Glycine, phenylalanine, and serine at pH = 5.9
  - Glycine, phenylalanine, and serine at pH = 6.0

## PROBLEM .....

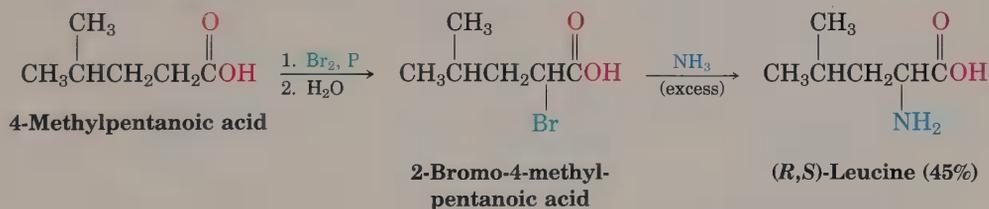
- 27.9 How can you account for the fact that tryptophan has a lower isoelectric point than histidine has, even though both have five-membered-ring nitrogen atoms? Which nitrogen in the five-membered ring of histidine is more basic?

## PROBLEM .....

- 27.10 Threonine has  $pK_{a1} = 2.1$  and  $pK_{a2} = 9.1$ . Use the Henderson–Hasselbalch equation to calculate the ratio of protonated and deprotonated forms at pH 1.5 and pH 10.0.

## 27.4 Synthesis of $\alpha$ -Amino Acids

$\alpha$ -Amino acids can be synthesized using some of the standard chemical reactions already discussed. One of the oldest methods of  $\alpha$ -amino-acid synthesis begins with alpha bromination of a carboxylic acid by treatment with bromine and phosphorus (the Hell–Volhard–Zelinskii reaction, Section 22.4). Nucleophilic substitution ( $S_N2$  reaction) of the  $\alpha$ -bromo acid with ammonia then yields an  $\alpha$ -amino acid.



Alternatively, higher product yields are obtained when the bromide displacement reaction is carried out by the Gabriel phthalimide method (Section 25.7), rather than by the ammonia method.

## PROBLEM .....

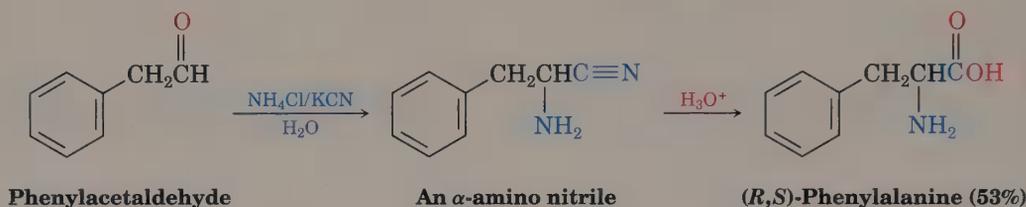
- 27.11 Show how you could prepare these  $\alpha$ -amino acids starting from the appropriate carboxylic acids:
- Phenylalanine
  - Valine

## PROBLEM.....

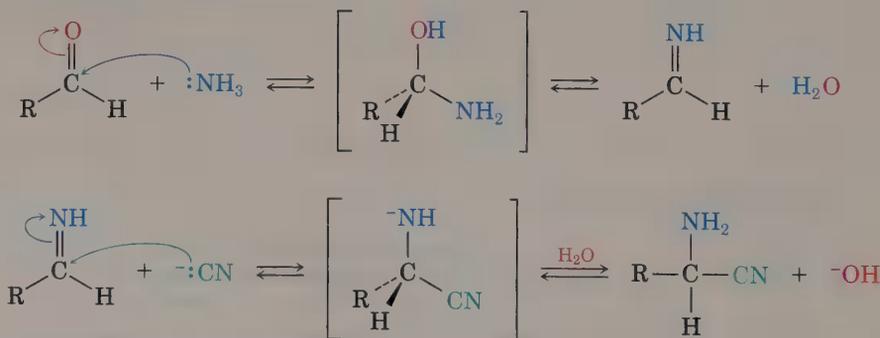
- 27.12 Show how a Gabriel amine synthesis (Section 25.7) might be used to prepare isoleucine.
- .....

## THE STRECKER SYNTHESIS

Another method for preparing racemic  $\alpha$ -amino acids is the **Strecker<sup>1</sup> synthesis**. Developed in 1850, this versatile two-step process involves treatment of an aldehyde with KCN and aqueous ammonia to yield an intermediate  $\alpha$ -amino nitrile. Hydrolysis of the nitrile then gives an  $\alpha$ -amino acid:



The first step of the Strecker synthesis, formation of an  $\alpha$ -amino nitrile by reaction of an aldehyde with ammonia and KCN, is simply a combination of two carbonyl-group reactions seen earlier in Chapter 19. Reaction of the aldehyde with ammonia yields an imine (Section 19.12), which then adds HCN in a nucleophilic addition step similar to that involved in cyanohydrin formation (Section 19.9). The mechanism is shown in Figure 27.3.

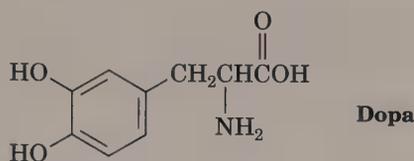


**Figure 27.3** Mechanism of  $\alpha$ -amino nitrile formation in the Strecker amino acid synthesis.

<sup>1</sup>Adolph Friedrich Ludwig Strecker (1822–1871); Ph.D. Giessen (1842); assistant to Liebig at Tübingen.

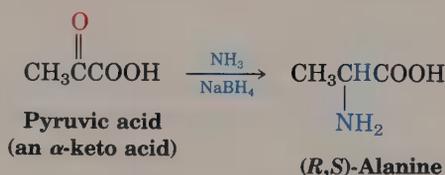
## PROBLEM.....

- 27.13 The rare amino acid L-dopa (3,4-dihydroxyphenylalanine) is useful as a drug against Parkinson's disease. Show how ( $\pm$ )-dopa might be synthesized from 3,4-dihydroxyphenylacetaldehyde.

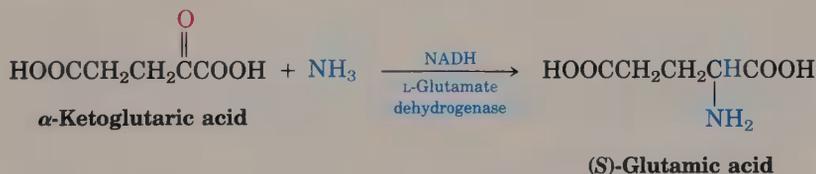


### REDUCTIVE AMINATION OF $\alpha$ -KETO ACIDS: BIOSYNTHESIS

Yet a third method for the synthesis of  $\alpha$ -amino acids is by reductive amination of an  $\alpha$ -keto acid with ammonia and a reducing agent (Section 25.7):



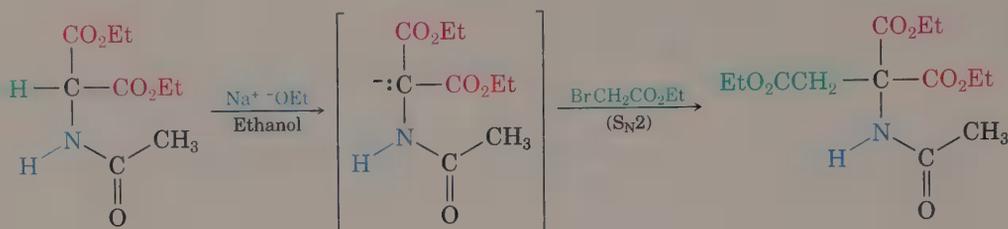
Although not widely used, this reductive amination method is interesting because it is a close laboratory analogy of a pathway by which some amino acids are biosynthesized in nature. For example, the major route for glutamic acid synthesis in most organisms is reductive amination of  $\alpha$ -ketoglutaric acid. The biological reducing agent is the rather complex molecule nicotinamide adenine dinucleotide (NADH), and the reaction is catalyzed by an enzyme, L-glutamate dehydrogenase. Nevertheless, the fundamental chemical principles of this biosynthetic reaction are identical to those of the laboratory reaction.



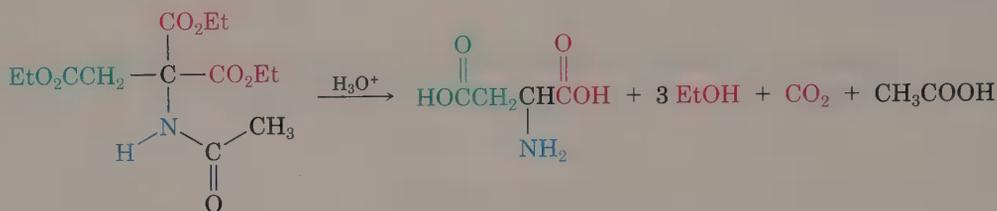
### THE AMIDOMALONATE SYNTHESIS

The most general method of preparation for  $\alpha$ -amino acids is the **amido-malonate synthesis**. This route, a straightforward extension of the malonic ester synthesis (Section 22.9), involves initial conversion of diethyl acetamidomalonate into its enolate anion by treatment with base, followed by  $\text{S}_{\text{N}}2$  reaction with a primary alkyl halide. Hydrolysis and decarboxylation

occur when the alkylated product is warmed with aqueous acid, and a racemic  $\alpha$ -amino acid results. For example, aspartic acid is prepared in good yield when diethyl acetamidomalonate is alkylated with ethyl bromoacetate, followed by hydrolysis and decarboxylation:



Diethyl acetamidomalonate

*(R,S)*-Aspartic acid (55%)

## PROBLEM.....

- 27.14 Show the alkyl halides you would use to prepare these  $\alpha$ -amino acids by the amidomalonate method.
- (a) Leucine                      (b) Histidine                      (c) Tryptophan                      (d) Methionine

## PROBLEM.....

- 27.15 Serine can be synthesized by a simple variation of the amidomalonate method using formaldehyde rather than an alkyl halide. Can you suggest how this might be done?
- .....

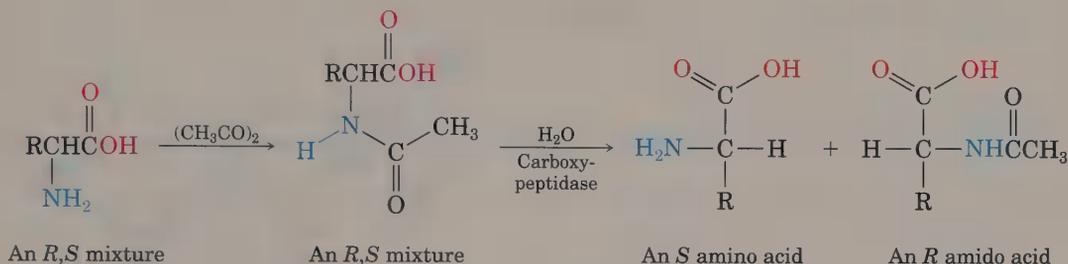
## 27.5 Resolution of *R,S* Amino Acids

The synthesis of chiral amino acids from achiral precursors by any one of the methods just described yields a racemic mixture—an equal mixture of *S* and *R* products. In order to use these synthetic amino acids for the laboratory synthesis of naturally occurring peptides, the racemic mixture must first be resolved into pure enantiomers.

Often this resolution can be done by the general method discussed earlier (Section 25.5), whereby the racemic mixture is converted into a mixture of diastereomeric salts by reaction with a chiral acid or base, and the different diastereomers are then separated by fractional crystallization.

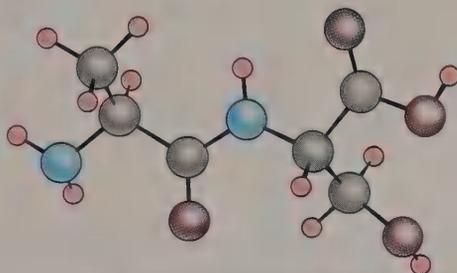
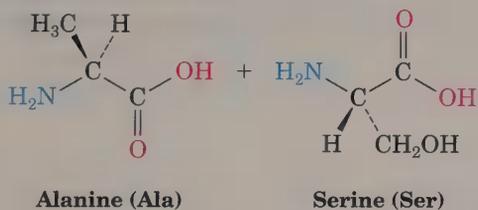
Alternatively, biological methods of resolution can be used. **Enzymes** are chiral biological catalysts that often show an astounding selectivity

toward one enantiomer of an *R,S* mixture. For example, the enzyme carboxypeptidase selectively catalyzes the hydrolysis of *S* amido acids but not *R* amido acids. We can therefore resolve an *R,S* mixture of amino acids by first allowing the mixture to react with acetic anhydride to form the *N*-acetyl derivatives. Selective hydrolysis of the *R,S* amido acid mixture with carboxypeptidase then yields a mixture of the desired *S* amino acid and the unchanged *N*-acetyl *R* amido acid, which can be separated by usual chemical techniques.



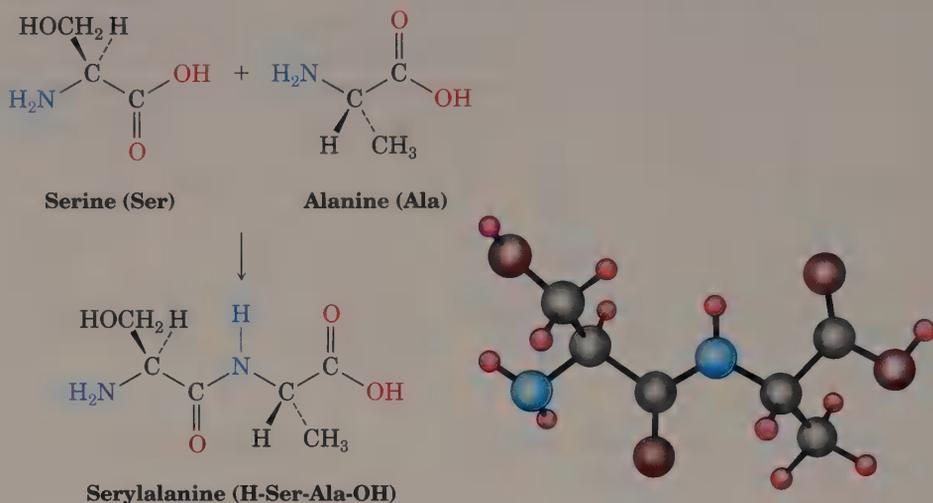
## 27.6 Peptides

**Peptides** are amino acid polymers in which the individual amino acid units, called **residues**, are linked together by amide, or peptide, bonds. An amino group from one residue forms an amide bond with the carboxyl of a second residue; the amino group of the second forms an amide bond with the carboxyl of a third, and so on. For example, alanylserine is the dipeptide formed when an amide bond is made between the alanine carboxyl and the serine amino group:



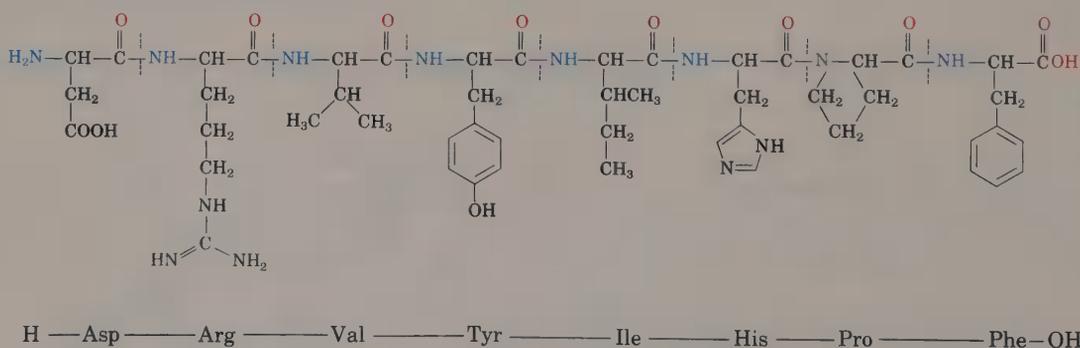
Note that two dipeptides can result from reaction between alanine and serine, depending on which carboxyl group reacts with which amino group.

If the alanine amino group reacts with the serine carboxyl, serylalanine results:



By convention, peptides are always written with the **N-terminal amino acid** (the one with the free  $-\text{NH}_2$  group) on the left and the **C-terminal amino acid** (the one with the free  $-\text{COOH}$  group) on the right. The name of the peptide is usually indicated by using the three-letter abbreviations listed in Table 27.1 for each amino acid. An H- is sometimes appended to the abbreviation of the leftmost amino acid to underscore its position as the N-terminal residue, and an -OH is sometimes appended to the abbreviation of the rightmost amino acid (C-terminal residue). Thus, alanylserine is abbreviated H-Ala-Ser-OH and serylalanine is abbreviated H-Ser-Ala-OH.

The number of possible isomeric peptides increases rapidly as the number of amino acid units increases. There are six ways in which three amino acids can be joined and more than 40,000 ways in which the eight amino acids present in the blood-pressure-regulating hormone angiotensin II can be joined (Figure 27.4).



**Figure 27.4** The structure of angiotensin II, a blood-pressure-regulating hormone present in blood plasma.

## PROBLEM .....

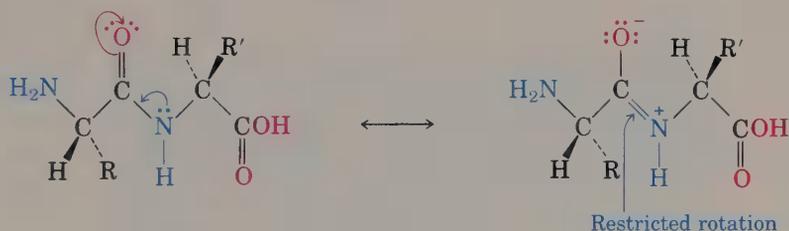
- 27.16 Name the six possible isomeric tripeptides that contain valine, tyrosine, and glycine. Use the three-letter shorthand notation for each amino acid.

## PROBLEM .....

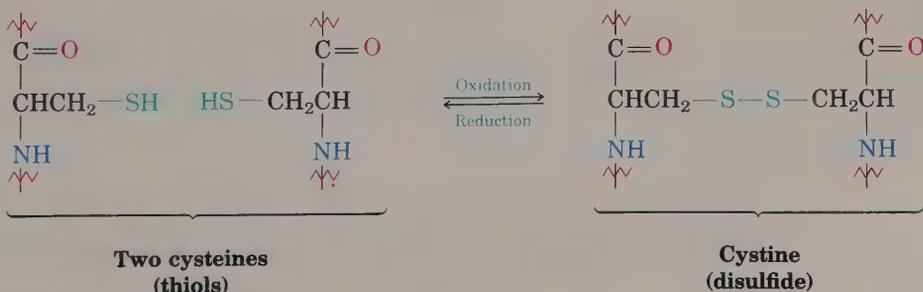
- 27.17 Draw the full structure of H-Met-Pro-Val-Gly-OH, and indicate where the amide bonds are.

## 27.7 Covalent Bonding in Peptides

Amide bonds in peptides are similar to the simple amide bonds we've already discussed (Section 25.4). Amide nitrogens are nonbasic because their unshared electron pair is delocalized by orbital overlap with the carbonyl group. This overlap imparts a certain amount of double-bond character to the amide C-N bond and restricts its rotation.

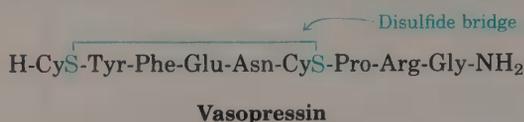


A second kind of covalent bonding in peptides occurs when a disulfide linkage, R-S-S-R, is formed between two cysteine residues. The linkage is sometimes represented by writing CyS with a capital "S" for sulfur and then drawing a line from one CyS to the other: CyS—CyS. As we saw in Section 17.12, disulfide bonds are easily formed by mild oxidation of thiols, RSH, and are easily cleaved back to thiols by mild reduction.



Disulfide bonds between cysteine residues in two different peptide chains link the otherwise separate chains together. Alternatively, a disulfide bond between two cysteine residues within the same chain causes a loop in the chain. Such is the case with the nonapeptide vasopressin, an antidiuretic

hormone involved in controlling water balance in the body. Note that the C-terminal end of vasopressin occurs as the primary amide,  $-\text{CONH}_2$ , rather than as the free acid.

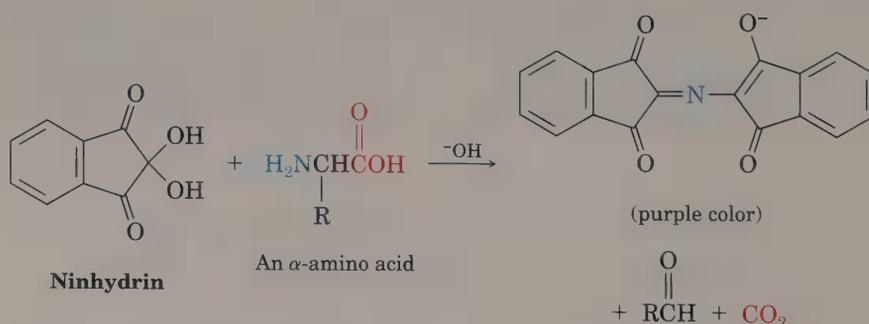


## 27.8 Peptide Structure Determination: Amino Acid Analysis

Determining the structure of a peptide requires finding the answers to three questions: What amino acids are present? How much of each is present? In what sequence do the amino acids occur in the peptide chain? The answers to the first two questions are provided by an **amino acid analyzer**.

An amino acid analyzer is an automated instrument based on analytical techniques worked out in the 1950s at the Rockefeller Institute by William Stein<sup>2</sup> and Stanford Moore.<sup>3</sup> The first step is to break the peptide down into its constituent amino acids by reducing all disulfide bonds and hydrolyzing all amide bonds with 6M HCl. Chromatography (Section 12.2) of the resultant amino acid mixture using a series of aqueous buffers as the mobile phase then effects a separation into component amino acids.

As each amino acid elutes from the end of the chromatography column, it mixes with a solution of *ninhydrin*, a reagent that reacts with  $\alpha$ -amino acids to form an intense purple color. The purple color is detected by a spectrometer, and a plot of elution time versus spectrometer absorbance is obtained.

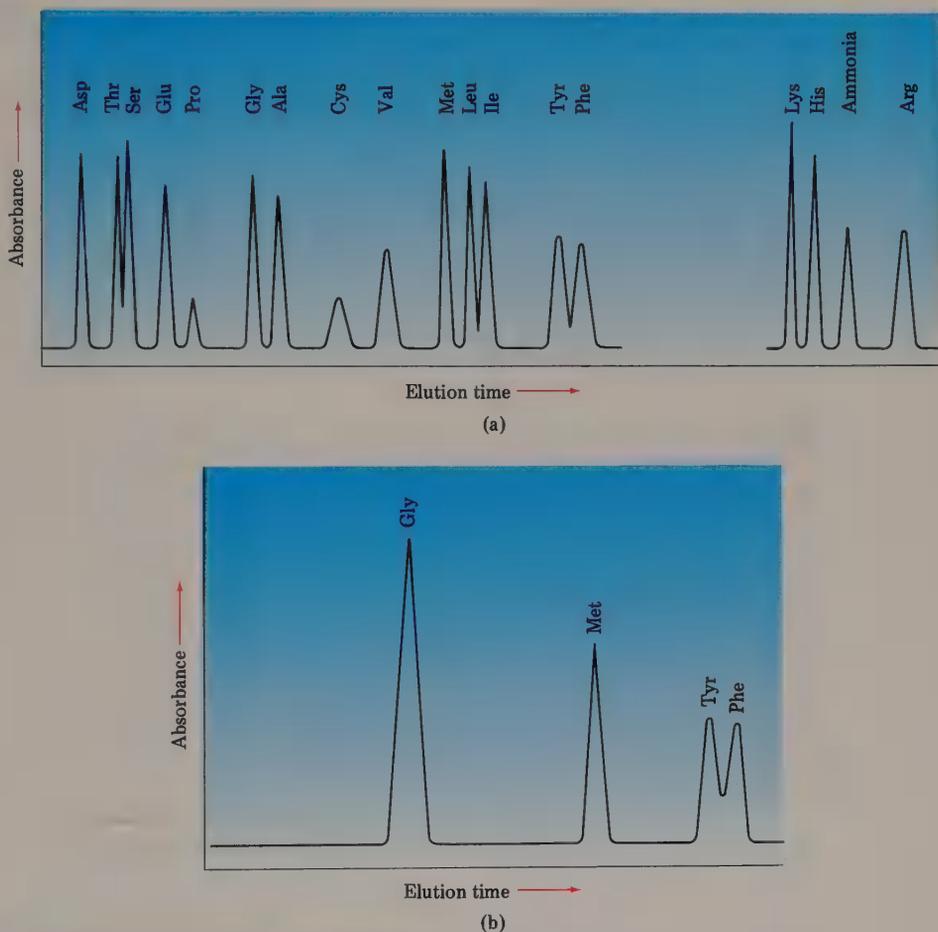


Since the amount of time required for a given amino acid to elute from the chromatography column is reproducible, the identity of all amino acids in a peptide of unknown composition can be determined simply by noting the various elution times. The amount of each amino acid in a sample is

<sup>2</sup>William H. Stein (1911–1980); b. New York; Ph.D. Columbia; professor, Rockefeller Institute; Nobel Prize (1972).

<sup>3</sup>Stanford Moore (1913–1982); b. Chicago; Ph.D. Wisconsin; professor, Rockefeller Institute; Nobel Prize (1972).

determined by measuring the intensity of the purple color resulting from its reaction with ninhydrin. Figure 27.5 shows the results of amino acid analysis of a standard equimolar mixture of 17  $\alpha$ -amino acids and compares them to results obtained from analysis of methionine enkephalin, a pentapeptide with morphine-like analgesic activity.



**Figure 27.5** Amino acid analysis of (a) an equimolar amino acid mixture and (b) methionine enkephalin (H-Tyr-Gly-Gly-Phe-Met-OH).

**PROBLEM** .....

**27.18** Show the structures of the products obtained on reaction of valine with ninhydrin.

**PROBLEM** .....

**27.19** The data for amino acid analysis in Figure 27.5(a) indicate that proline is not easily detected by reaction with ninhydrin; only a very small peak is seen on the chromatogram. Explain.

.....

## 27.9 Peptide Sequencing: The Edman Degradation

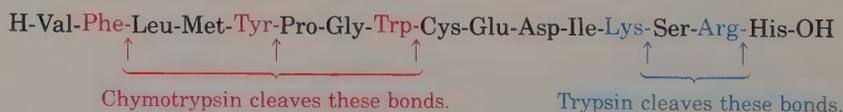
With the identity and amount of each amino acid known, the final task of structure determination is to *sequence* the peptide—that is, to find out in what order the amino acids are linked together. The general idea of peptide sequencing is to cleave one amino acid residue at a time from the end of the peptide chain (either N terminus or C terminus). That terminal amino acid is then separated and identified, and the cleavage reactions are repeated on the chain-shortened peptide until the entire peptide sequence is determined.

Most peptide sequencing is now done by **Edman<sup>4</sup> degradation**, an efficient method of N-terminal analysis. Automated Edman protein sequencers are available that allow a series of 20 or more repetitive sequencing steps to be carried out before a buildup of unwanted by-products begins to interfere with the results.

Edman degradation involves treatment of a peptide with phenyl isothiocyanate,  $C_6H_5-N=C=S$ , followed by mild acid hydrolysis, as shown in Figure 27.6. The first step attaches a marker to the  $-NH_2$  group of the N-terminal amino acid residue, and the second step splits the N-terminal residue from the chain, yielding a *phenylthiohydantoin* derivative plus the chain-shortened peptide. The phenylthiohydantoin is then identified chromatographically by comparison with known derivatives of the common amino acids, and the chain-shortened peptide is resubmitted to another round of Edman degradation.

Complete sequencing of large peptides and proteins by Edman degradation is impractical since buildup of unwanted by-products limits the method to about 25 cycles. Instead, a large peptide chain is first cleaved by partial hydrolysis into a number of smaller fragments, the sequence of each fragment is determined, and the individual fragments are fitted together like pieces in a jigsaw puzzle.

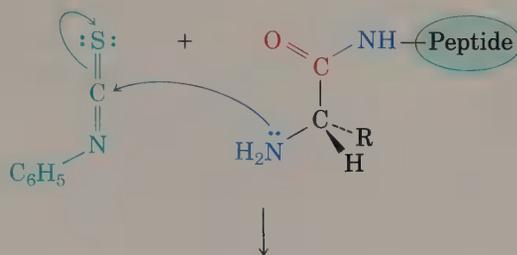
Partial hydrolysis of a peptide can be carried out either chemically with aqueous acid or enzymatically with enzymes such as trypsin and chymotrypsin. Acidic hydrolysis is unselective and leads to a more or less random mixture of small fragments. Enzymic hydrolysis, however, is quite specific. Trypsin catalyzes hydrolysis only at the carboxyl side of the basic amino acids arginine and lysine; chymotrypsin cleaves only at the carboxyl side of the aryl-substituted amino acids phenylalanine, tyrosine, and tryptophan.



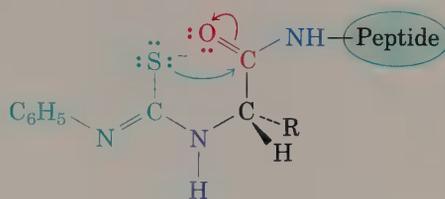
To take an example of peptide sequencing, let's look at a hypothetical structure determination of angiotensin II, a hormonal octapeptide involved

<sup>4</sup>Pebr Edman (1916– ); b. Stockholm; M.D. Karolinska Institute (E. Jorpes); professor, University of Lund.

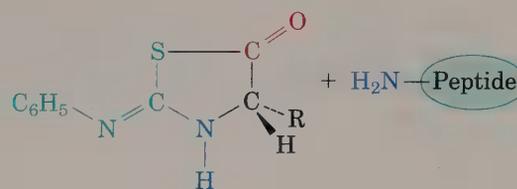
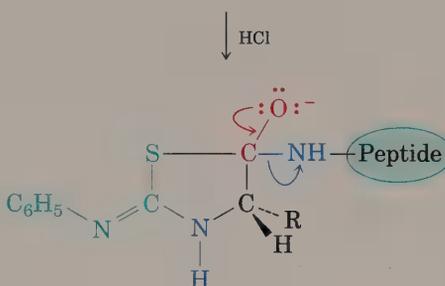
Nucleophilic addition of the peptide terminal amino group to phenylisothiocyanate yields an *N*-phenylthiourea derivative.



Acid-catalyzed cyclization then yields a tetrahedral intermediate . . .

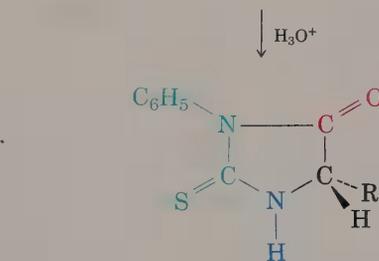


. . . which expels the chain-shortened peptide and forms a thiazolinone.



The thiazolinone rearranges in the presence of aqueous acid to yield the final *N*-phenylthiohydantoin derivative.

A thiazolinone



An *N*-phenylthiohydantoin

**Figure 27.6** Mechanism of the Edman degradation for *N*-terminal analysis of peptides.

in controlling blood pressure by regulating the sodium–potassium salt balance in the body.

1. Amino acid analysis of angiotensin II shows the presence of eight different amino acids: Arg, Asp, His, Ile, Phe, Pro, Tyr, and Val in equimolar amounts.
2. An N-terminal analysis by the Edman method shows that angiotensin II has an aspartic acid residue at the N terminus.
3. Partial hydrolysis of angiotensin II with dilute hydrochloric acid might yield the following fragments, whose sequences could be determined by Edman degradation:
  - a. H-Asp-Arg-Val-OH
  - b. H-Ile-His-Pro-OH
  - c. H-Arg-Val-Tyr-OH
  - d. H-Pro-Phe-OH
  - e. H-Val-Tyr-Ile-OH
4. Matching the overlapping regions of the various fragments provides the full sequence of angiotensin II:
  - a. H-Asp-Arg-Val-OH
  - c.       H-Arg-Val-Tyr-OH
  - e.               H-Val-Tyr-Ile-OH
  - b.                       H-Ile-His-Pro-OH
  - d.                               H-Pro-Phe-OH

H-Asp-Arg-Val-Tyr-Ile-His-Pro-Phe-OH

#### Angiotensin II

The structure of angiotensin II is relatively simple—the entire sequence could easily be done by a protein sequenator—but the methods and logic used here are the same as those used to solve more complex structures. Indeed, single protein chains with more than 400 amino acids have been sequenced by these methods.

#### PROBLEM .....

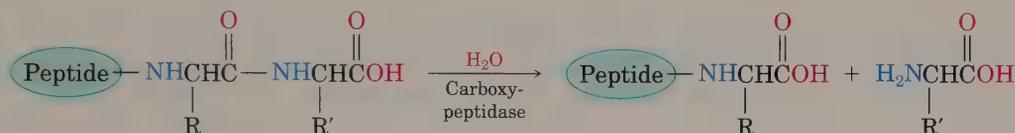
- 27.20** What fragments would result if angiotensin II were cleaved with trypsin? With chymotrypsin?

#### PROBLEM .....

- 27.21** Give the amino acid sequence of hexapeptides that produce these fragments on partial acid hydrolysis:
- (a) Arg, Gly, Ile, Leu, Pro, Val gives H-Pro-Leu-Gly-OH, H-Arg-Pro-OH, H-Gly-Ile-Val-OH
  - (b) Asp, Leu, Met, Trp, Val<sub>2</sub> gives H-Val-Leu-OH, H-Val-Met-Trp-OH, H-Trp-Asp-Val-OH
- .....

## 27.10 Peptide Sequencing: C-Terminal Residue Determination

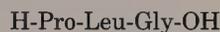
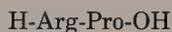
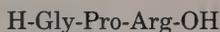
The Edman degradation is an excellent method of analysis for the N-terminal residue, but a complementary method of analysis for the C-terminal residue is also valuable. The best method currently available makes use of the enzyme carboxypeptidase to specifically cleave the C-terminal amide bond in a peptide chain:



The analysis is carried out by incubating the polypeptide with carboxypeptidase and watching for the appearance of the first free amino acid that appears in solution. Of course, further degradation also occurs, since a new C terminus is produced when the first amino acid residue is cleaved off, until ultimately the entire peptide is hydrolyzed.

### PROBLEM .....

- 27.22** A hexapeptide with the composition Arg, Gly, Leu, Pro<sub>3</sub> is found to have proline at both C-terminal and N-terminal positions. Partial hydrolysis gives the following fragments:



What is the structure of the hexapeptide?

### PROBLEM .....

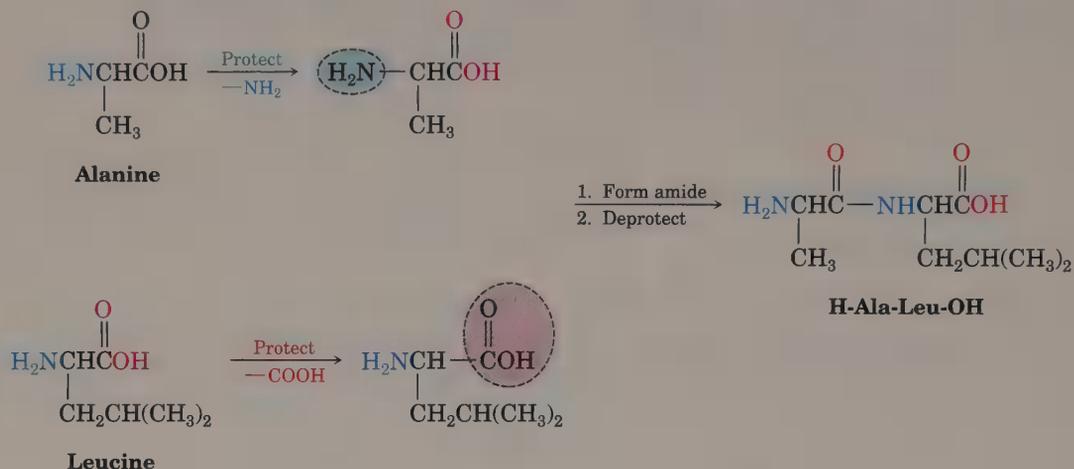
- 27.23** Propose two structures for a tripeptide that gives Leu, Ala, and Phe on hydrolysis but does not react with carboxypeptidase and does not react with phenyl isothiocyanate.
- .....

## 27.11 Peptide Synthesis

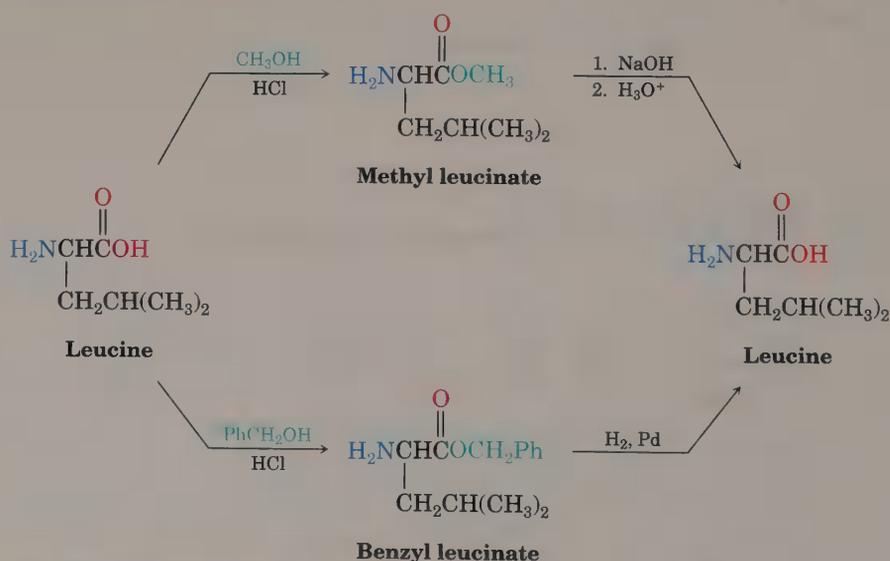
Once a peptide's structure has been determined, synthesis is often the next goal. This might be done either as a final proof of structure or as a means of obtaining larger amounts of a valuable peptide for biological evaluation. Although ordinary amide bonds are usually formed by reaction between amines and acid chlorides (Section 21.8), peptide synthesis is much more complex because of the requirement for specificity. Many different amide links must be formed, and they must be formed in a specific order rather than at random. We can't expect simply to place a mixture of amino acids in a flask and obtain a single polypeptide product.

The solution to the specificity problem is *protection*. We can force a reaction to take only the desired course by protecting all of the amine and

acid functional groups except for those we want to have react. For example, if we want to couple alanine with leucine to synthesize H-Ala-Leu-OH, we can protect the amino group of alanine and the carboxyl group of leucine to render them unreactive, then form the desired amide bond, and then remove the protecting groups.



Many different amino- and carboxyl-protecting groups have been devised but only a few are widely used. Carboxyl groups are often protected simply by converting them into methyl or benzyl esters. These groups are easily introduced by standard methods of ester formation and are easily removed by mild hydrolysis with aqueous sodium hydroxide. As Figure 27.7 shows, benzyl esters can also be cleaved by catalytic hydrogenolysis of the weak benzylic C–O bond ( $\text{ArCH}_2\text{-OCOR} + \text{H}_2 \rightarrow \text{ArCH}_3 + \text{RCOOH}$ ).



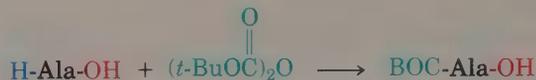
**Figure 27.7** Protection of an amino acid carboxyl group by ester formation.



carboxylic acid group into a reactive acylating agent that then undergoes a further nucleophilic acyl substitution with the amine.

The five steps needed to synthesize H-Ala-Leu-OH are summarized:

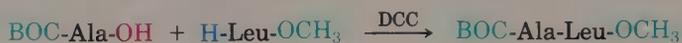
*Step 1* Protect the amino group of alanine as the BOC derivative:



*Step 2* Protect the carboxyl group of leucine as the methyl ester:



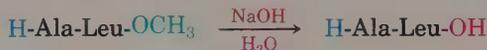
*Step 3* Couple the two protected amino acids using DCC:



*Step 4* Remove the BOC protecting group by acid treatment:



*Step 5* Remove the methyl ester by basic hydrolysis:



These steps can be repeated to add one amino acid at a time to the growing chain or to link two peptide chains together. Many remarkable achievements in peptide synthesis have been reported, including a complete synthesis of human insulin. Insulin, whose structure is shown in Figure 27.9, is composed of two chains totaling 51 amino acids linked by two disulfide bridges. Its structure was determined by Frederick Sanger,<sup>5</sup> who received the 1958 Nobel Prize for his work.

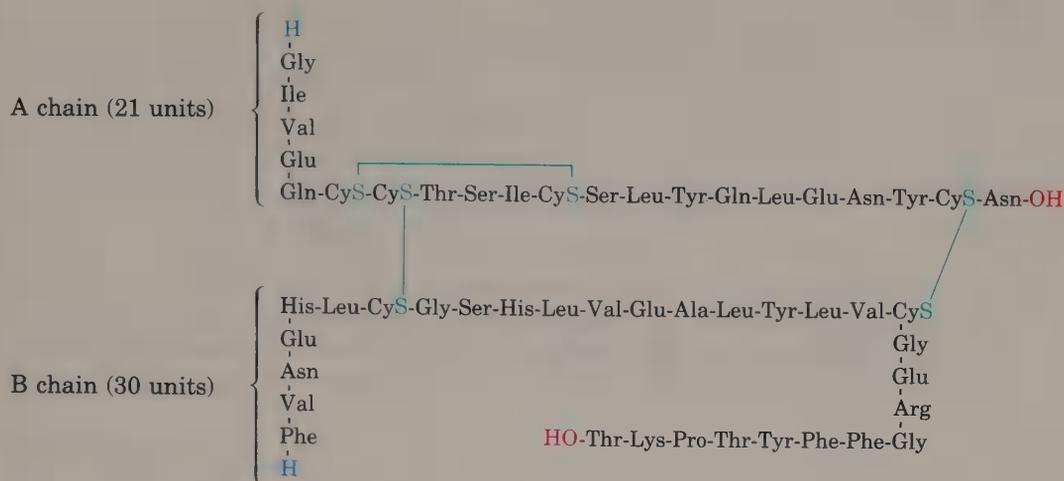
#### PROBLEM .....

- 27.24** Show the mechanism for formation of a BOC derivative by reaction of an amino acid with di-*tert*-butyl dicarbonate.

#### PROBLEM .....

- 27.25** Write all five steps required for the synthesis of H-Leu-Ala-OH from alanine and leucine.

<sup>5</sup>Frederick Sanger (1918– ); b. Gloucestershire, England; Ph.D. Cambridge; professor, Cambridge University; Nobel Prize (1958, 1980).



**Figure 27.9** Structure of human insulin. Two separate chains totaling 51 amino acids are linked by two disulfide bridges.

**PROBLEM** .....

**27.26** How would you prepare these tripeptides?

(a) H-Leu-Ala-Gly-OH

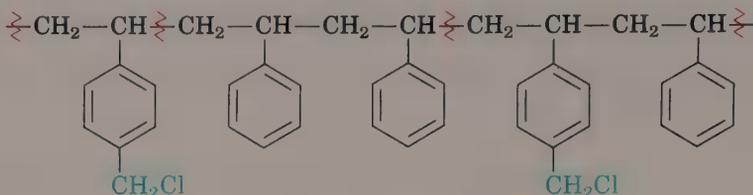
(b) H-Gly-Leu-Ala-OH

## 27.12 Automated Peptide Synthesis: The Merrifield Solid-Phase Technique

The synthesis of large peptide chains by sequential addition of one amino acid at a time is a long and arduous task. An immense simplification is possible, however, using the **solid-phase synthesis method** introduced by R. Bruce Merrifield<sup>6</sup> at the Rockefeller University. In the Merrifield method, peptide synthesis is carried out on solid polymer beads of polystyrene, pre-

<sup>6</sup>Robert Bruce Merrifield (1921– ); Ph.D. University of California, Los Angeles (1949); professor, Rockefeller University; Nobel Prize (1984).

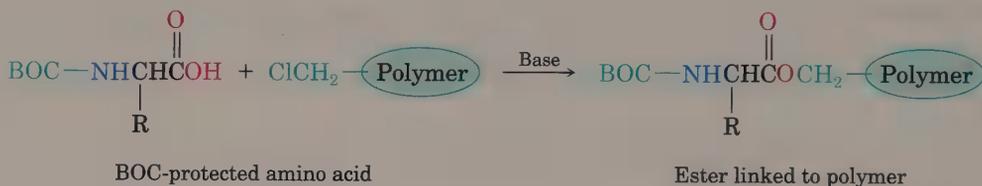
pared so that one of every 100 or so benzene rings bears a chloromethyl ( $-\text{CH}_2\text{Cl}$ ) group:



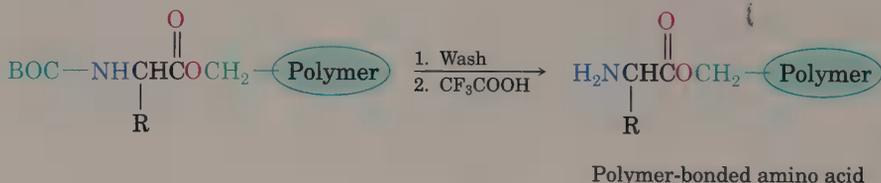
**Chloromethylated polystyrene**

In the standard solution-phase method discussed in the previous section, a methyl ester was used to protect the carboxyl group during formation of the amide bond. In the solid-phase method, however, the solid *polymer* is the ester protecting group. Four steps are required in solid-phase peptide synthesis:

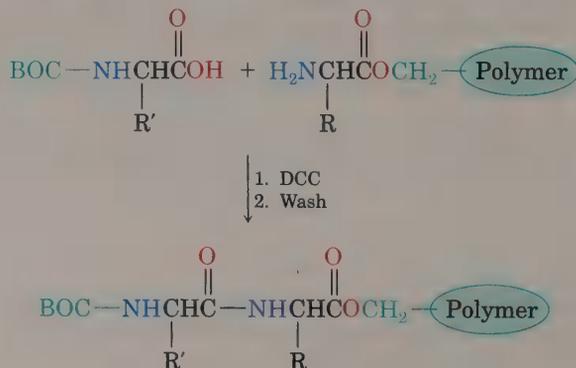
**Step 1** A covalent ester linkage is formed by reaction between a BOC-protected amino acid and the chloromethyl groups on the polystyrene polymer:



**Step 2** After formation of the ester linkage is complete, the insoluble, polymer-bonded amino acid is washed free of excess reagents and treated with trifluoroacetic acid to remove the BOC group:

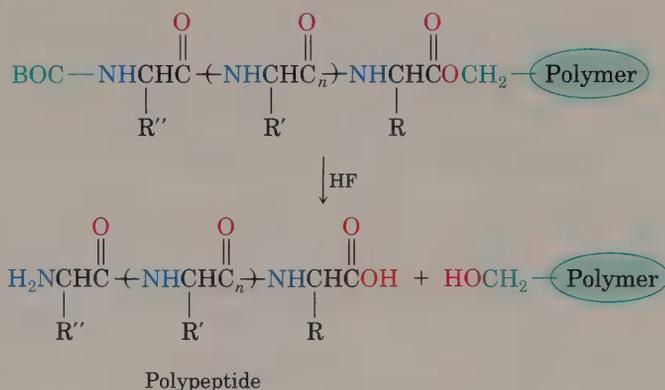


**Step 3** A second BOC-protected amino acid is added along with the coupling reagent, DCC. Peptide-bond formation occurs, and excess reagents are removed by washing the insoluble polymer:



Step 2 is repeated to again remove a BOC group, and step 3 is repeated to add a third amino acid unit to the chain. In this way, dozens or even a hundred amino acid units can be efficiently and specifically linked to synthesize the desired polymer-bonded peptide:

*Step 4* After the proper number of coupling steps have been done and the desired peptide has been made, treatment with anhydrous hydrogen fluoride cleaves the ester bond to the polymer, yielding free peptide:



The solid-phase technique has been automated, and computer-controlled peptide synthesizers are available for automatically repeating the coupling and deprotection steps with different amino acids as many times as desired. Each step occurs in extremely high yield, and mechanical losses are minimized since the peptide intermediates are never removed from the insoluble polymer until the final step. Among the many remarkable achievements recorded by Merrifield is the synthesis of bovine pancreatic ribonuclease, a protein containing 124 amino acid units. The entire synthesis required only 6 weeks and took place in 17% overall yield.

## 27.13 Classification of Proteins

Proteins are classified into two major types according to their composition. **Simple proteins**, such as blood serum albumin, are those that yield only amino acids and no other compounds on hydrolysis. **Conjugated proteins**, which are much more common than simple proteins, yield other compounds in addition to amino acids on hydrolysis.

Conjugated proteins can be further classified according to the chemical nature of the non-amino acid portion. **Glycoproteins** contain a carbohydrate part, **lipoproteins** contain a fatty part, **nucleoproteins** contain a nucleic acid part, and so on (Table 27.2). Glycoproteins are particularly widespread in nature and make up a large part of the membrane coating around living cells.

Another way to classify proteins is as *fibrous* or *globular*, according to their three-dimensional shape. **Fibrous proteins**, such as collagen and keratin, consist of polypeptide chains arranged side by side in long filaments. Because these proteins are tough and insoluble in water, they're used in

nature for structural materials such as tendons, hooves, horns, and muscles. **Globular proteins**, by contrast, are usually coiled into compact, nearly spherical shapes. These proteins are generally soluble in water and are mobile within cells. Most of the 2000 or so known enzymes, as well as hormonal and transport proteins, are globular. Table 27.3 lists some common examples of both fibrous and globular proteins.

**TABLE 27.2** Classification of Some Conjugated Proteins

<i>Class</i>	<i>Non-amino acid group</i>	<i>Weight of non-amino acid (%)</i>
<b>Glycoproteins</b>		
$\gamma$ -Globulin	Carbohydrate	10
Carboxypeptidase Y	Carbohydrate	17
Interferon	Carbohydrate	20
<b>Lipoproteins</b>		
Plasma $\beta$ lipoprotein	Fats, cholesterol	80
<b>Nucleoproteins</b>		
Ribosomal proteins	Ribonucleic acid	60
Tobacco mosaic virus	Ribonucleic acid	5
<b>Phosphoproteins</b>		
Casein	Phosphate esters	4
<b>Metalloproteins</b>		
Ferritin	Iron oxide	23
Hemoglobin	Iron	0.3

**TABLE 27.3** Conformational Classes of Proteins

<i>Protein</i>	<i>Description</i>
<b>Fibrous proteins (insoluble)</b>	
Collagen	Connective tissue, tendons
$\alpha$ -Keratin	Hair, horn, skin, nails
Elastin	Elastic connective tissue
<b>Globular proteins (soluble)</b>	
Insulin	Hormone controlling glucose metabolism
Lysozyme	Hydrolytic enzyme
Ribonuclease	Enzyme controlling RNA synthesis
Albumins	Proteins coagulated by heat
Immunoglobulins	Proteins involved in immune response
Myoglobin	Protein involved in oxygen transport

Yet a third way to classify proteins is according to function. As shown in Table 27.4, there is an extraordinary diversity to the biological roles of proteins.

**TABLE 27.4** Some Biological Functions of Proteins

<i>Type</i>	<i>Function and example</i>
Enzymes	Proteins such as chymotrypsin that act as biological catalysts
Hormones	Proteins such as insulin that regulate body processes
Protective proteins	Proteins such as antibodies that fight infection
Storage proteins	Proteins such as casein that store nutrients
Structural proteins	Proteins such as keratin, elastin, and collagen that form an organism's structure
Transport proteins	Proteins such as hemoglobin that transport oxygen and other substances through the body

## 27.14 Protein Structure

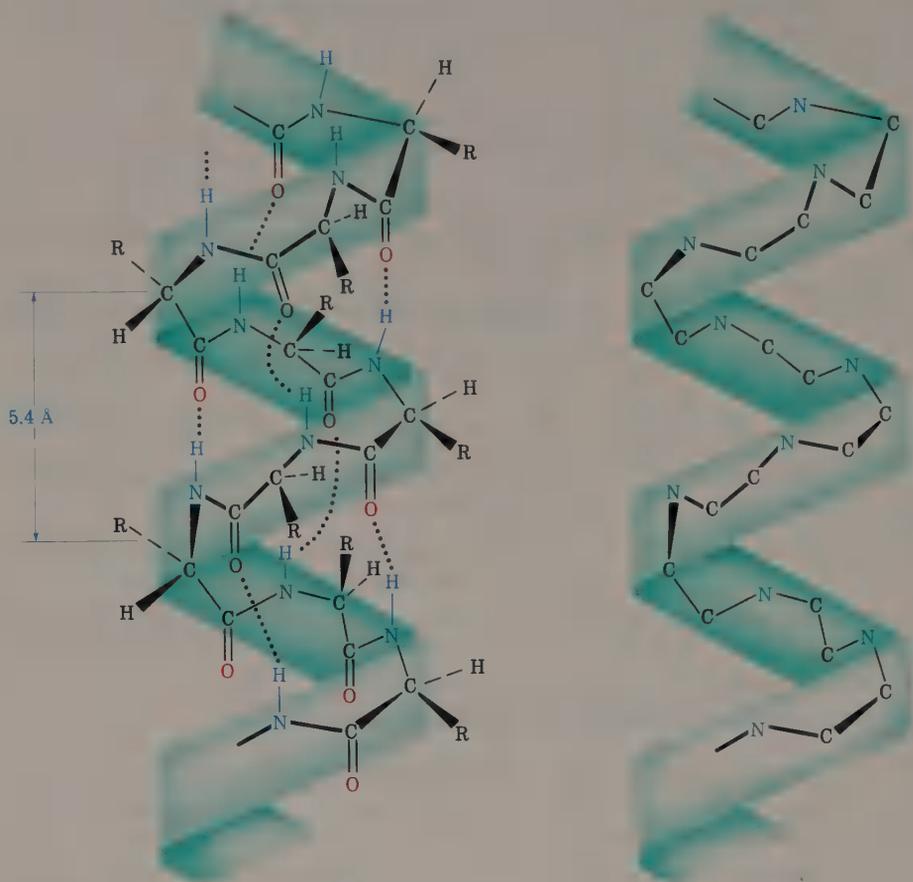
Proteins are so large in comparison to simple organic molecules that the word *structure* takes on a broader meaning when applied to these immense macromolecules than it does with other organic compounds. In fact, chemists speak of four different levels of structure when describing proteins. At its simplest, protein structure is the sequence in which amino acid residues are bound together. Called the **primary structure** of a protein, this is the most fundamental structural level.

There is much more to protein structure than just amino acid sequence though. The chemical properties of a protein are also dependent on higher levels of structure—on exactly how the peptide backbone is folded to give the molecule a specific three-dimensional shape. Thus, the term **secondary structure** refers to the way in which segments of the peptide backbone are oriented into a regular pattern; **tertiary structure** refers to the way in which the entire protein molecule is coiled into an overall three-dimensional shape; and **quaternary structure** refers to the way in which several protein molecules come together to yield large aggregate structures.

Let's look at three examples— $\alpha$ -keratin (fibrous), fibroin (fibrous), and myoglobin (globular)—to see how higher structure affects protein properties.

### $\alpha$ -KERATIN

$\alpha$ -Keratin is the fibrous structural protein found in wool, hair, nails, and feathers. Studies show that  $\alpha$ -keratin is coiled into a right-handed secondary structure as illustrated in Figure 27.10. This so-called  **$\alpha$ -helix** is stabilized



**Figure 27.10** The helical secondary structure of  $\alpha$ -keratin.

by hydrogen bonding between amide N–H groups and other amide carbonyl groups four residues away. Although the strength of a single hydrogen bond (about 5 kcal/mol) is only about 5% the strength of a C–C or C–H covalent bond, the large number of hydrogen bonds made possible by helical winding imparts a great deal of stability to the  $\alpha$ -helical structure. Each coil of the helix (the *repeat distance*) contains 3.6 amino acid residues; the distance between coils is 5.40 Å.

Further evidence shows that the  $\alpha$ -keratins of wool and hair also have a definite quaternary structure. The individual helices are themselves coiled about one another to form a *superhelix* that accounts for the threadlike properties and strength of these proteins.

Although  $\alpha$ -keratin is the best example of an almost entirely helical protein, most globular proteins contain  $\alpha$ -helical *segments*. Both hemoglobin and myoglobin, for example, contain many short helical sections in their chains.

## FIBROIN

Fibroin, the fibrous protein found in silk, has a secondary structure called a  $\beta$ -pleated sheet. In this pleated-sheet structure, polypeptide chains line up in a parallel arrangement held together by hydrogen bonds between chains (Figure 27.11). Although not as common as the  $\alpha$ -helix, small  $\beta$ -pleated-sheet regions are often found in proteins where sections of peptide chains double back on themselves.

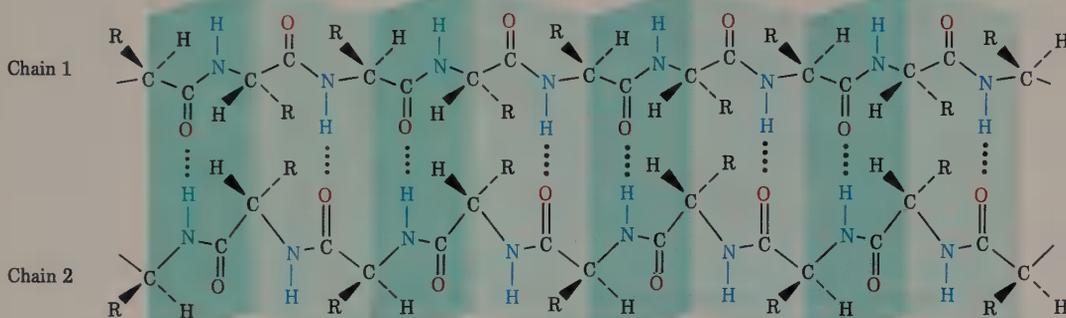


Figure 27.11 The  $\beta$ -pleated-sheet structure in silk fibroin.

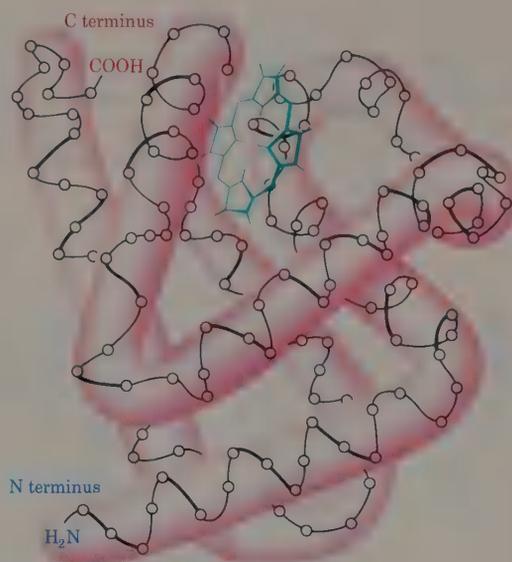
## MYOGLOBIN

Myoglobin is a small globular protein containing 153 amino acid residues in a single chain. A relative of hemoglobin, myoglobin is found in the skeletal muscles of sea mammals, where it stores oxygen needed to sustain the animals during long dives. X-ray evidence obtained by Sir John Kendrew<sup>7</sup> and Max Perutz<sup>8</sup> has shown that myoglobin consists of eight straight segments, each of which adopts an  $\alpha$ -helical secondary structure. These helical sections are connected by bends to form a compact, nearly spherical, tertiary structure (Figure 27.12).

Why does myoglobin adopt the shape it does? The forces that determine the tertiary structure of myoglobin and other globular proteins are the same simple forces that act on all molecules, regardless of size. By bending and twisting in exactly the right way, myoglobin achieves maximum stability. Although the bends appear irregular and the three-dimensional structure appears random, this isn't the case. All myoglobin molecules adopt this same shape because it is the most stable conformation.

<sup>7</sup>Sir John C. Kendrew (1917– ); b. Oxford, England; Ph.D. Cambridge; professor, Cambridge University; Nobel Prize (1962).

<sup>8</sup>Max Ferdinand Perutz (1914– ); b. Vienna; universities of Vienna and Cambridge; professor, Cambridge University; Nobel Prize (1962).

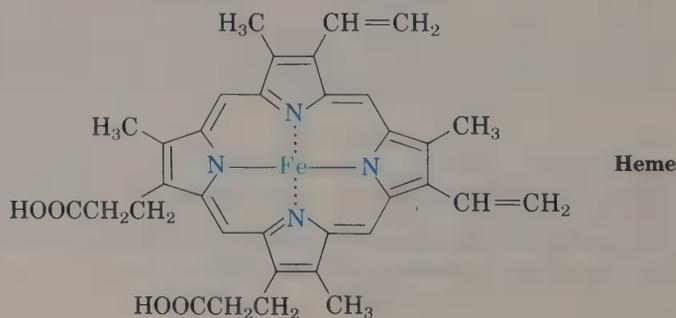


**Figure 27.12** Secondary and tertiary structure of myoglobin, a globular protein.

The most important forces stabilizing a protein's tertiary structure are the hydrophobic (water-repelling) interactions of hydrocarbon side chains on neutral amino acids. Those amino acids with neutral, nonpolar side chains have a strong tendency to congregate on the hydrocarbon-like interior of a protein molecule, away from the aqueous medium. Those acidic or basic amino acids with charged side chains, by contrast, tend to congregate on the exterior of the protein where they can be solvated by water.

Also important for stabilizing a protein's tertiary structure are the formation of disulfide bridges between cysteine residues, the formation of hydrogen bonds between nearby amino acids, and the development of ionic attractions, called *salt bridges*, between positively and negatively charged sites on the protein.

Note that myoglobin is a conjugated protein that contains a covalently bound organic group (a **prosthetic group**) called *heme*. A great many proteins contain such prosthetic groups, which are crucial to their mechanism of action.



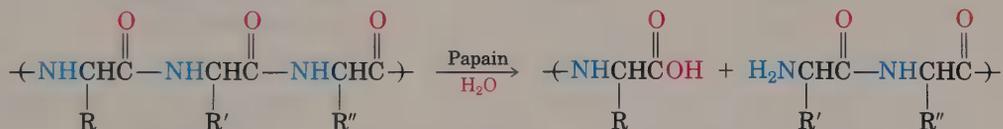
## PROBLEM.....

- 27.27 How can you account for the fact that proline is never present in a protein  $\alpha$ -helix? The  $\alpha$ -helical parts of myoglobin and other proteins stop whenever a proline residue is encountered in the chain.

## 27.15 Enzymes

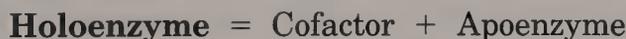
Enzymes, as we've seen, are large proteins that act as catalysts for biological reactions. Unlike many of the simple catalysts that chemists use in the laboratory, though, enzymes are usually specific in their action. Often, in fact, an enzyme can catalyze only a single reaction of a single compound, called the enzyme's *substrate*. For example, the enzyme amylase found in the human digestive tract catalyzes only the hydrolysis of starch to yield glucose; cellulose and other polysaccharides are untouched by amylase.

Different enzymes have different specificities. Some, such as amylase, are specific for a single substrate, but others operate on a range of substrates. Papain, for instance, a globular protein of 212 amino acids isolated from papaya fruit, catalyzes the hydrolysis of many kinds of peptide bonds. In fact, it's this ability to hydrolyze peptide bonds that makes papain useful as a meat tenderizer and a contact-lens cleaner.



Like all catalysts, enzymes don't affect the equilibrium constant of a reaction and can't bring about chemical changes that are otherwise unfavorable. Enzymes act only to lower the activation energy for reaction, thereby making the reaction take place faster or at lower temperatures. Starch and water, for example, react very slowly in the absence of a catalyst because the activation energy is too high. When amylase is present, though, the energy barrier is lowered, and the hydrolysis reaction occurs rapidly.

All of the 2000 or so known enzymes are globular proteins. In addition to the protein part, most enzymes also have small nonprotein parts called **cofactors**. The protein part in such enzymes is called an **apoenzyme**, and the combination of apoenzyme plus cofactor is called a **holoenzyme**. Only holoenzymes have biological activity; neither cofactor nor apoenzyme catalyzes reactions by itself.



Cofactors can be either inorganic ions such as  $\text{Zn}^{2+}$  or small organic molecules called **coenzymes**. The requirement of many enzymes for inorganic cofactors is the main reason for our dietary need of trace minerals. Iron, zinc, copper, manganese, and numerous other metal ions are all essential minerals that act as enzyme cofactors, though the exact biological role isn't known in all cases.

A variety of organic molecules act as coenzymes. Many, though not all, coenzymes are **vitamins**, small organic molecules that must be obtained in the diet and are required in trace amounts for proper growth. Table 27.5 lists the 13 known vitamins required in the human diet and their enzyme functions.

**TABLE 27.5** Vitamins and Their Enzyme Functions

<i>Vitamin</i>	<i>Enzyme function</i>	<i>Deficiency symptom</i>
<b>Water-Soluble Vitamins</b>		
Ascorbic acid (Vitamin C)	Hydroxylases	Bleeding gums, bruising
Thiamin (Vitamin B <sub>1</sub> )	Reductases	Fatigue, depression
Riboflavin (Vitamin B <sub>2</sub> )	Reductases	Cracked lips, scaly skin
Pyridoxine (Vitamin B <sub>6</sub> )	Aminotransferases	Anemia, irritability
Niacin	Reductases	Dermatitis, dementia
Folic acid (Vitamin M)	Methyltransferases	Megaloblastic anemia
Vitamin B <sub>12</sub>	Isomerases	Megaloblastic anemia, neurodegeneration
Pantothenic acid	Acyltransferases	Weight loss, irritability
Biotin (Vitamin H)	Carboxylases	Dermatitis, anorexia, depression
<b>Fat-Soluble Vitamins</b>		
Vitamin A	Visual system	Night blindness, dry skin
Vitamin D	Calcium metabolism	Rickets, osteomalacia
Vitamin E	Antioxidant	Hemolysis of red blood cells
Vitamin K	Blood clotting	Hemorrhage, delayed blood clotting

Enzymes are grouped into six main classes according to the kind of reaction they catalyze (Table 27.6). *Hydrolases* catalyze the hydrolysis of substrates; *isomerases* catalyze the isomerization of substrates; *ligases* catalyze the bonding together of two substrates with participation of ATP; *lyases* catalyze the breaking away of a small molecule such as H<sub>2</sub>O from a substrate or the reverse reaction; *oxidoreductases* catalyze oxidations and reductions of substrate molecules; and *transferases* catalyze the transfer of a group from one substrate to another.

Although some enzymes, like papain and trypsin, have uninformative common names, the systematic name of an enzyme has two parts, ending with *-ase*. The first part identifies the enzyme's substrate, and the second

TABLE 27.6 Classification of Enzymes

Main class	Some subclasses	Type of reaction catalyzed
Hydrolases	Lipases	Hydrolysis of an ester group
	Nucleases	Hydrolysis of a phosphate group
	Proteases	Hydrolysis of an amide group
Isomerases	Epimerases	Isomerization of stereogenic center
Ligases	Carboxylases	Addition of CO <sub>2</sub>
	Synthetases	Formation of new bond
Lyases	Decarboxylases	Loss of CO <sub>2</sub>
	Dehydrases	Loss of H <sub>2</sub> O
Oxidoreductases	Dehydrogenases	Introduction of double bond by removal of H <sub>2</sub>
	Oxidases	Oxidation
	Reductases	Reduction
Transferases	Kinases	Transfer of a phosphate group
	Transaminases	Transfer of an amino group

part identifies its class. For example, *hexose kinase* is an enzyme that catalyzes the transfer of a phosphate group from adenosine triphosphate (ATP) to glucose.

PROBLEM.....

27.28 To what classes do these enzymes belong?

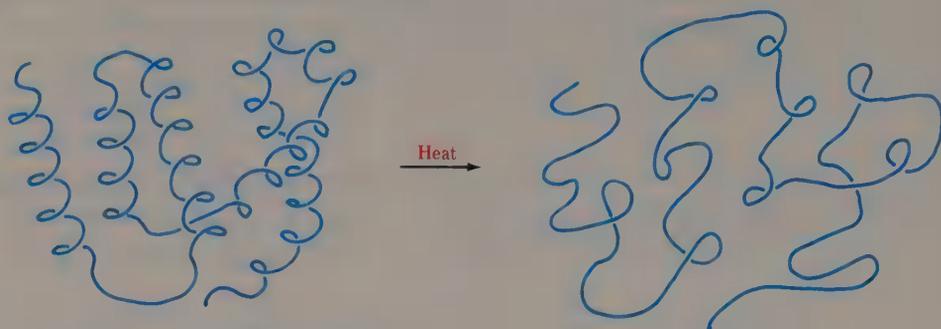
- (a) Pyruvate decarboxylase    (b) Chymotrypsin    (c) Alcohol dehydrogenase
- .....

## 27.16 Protein Denaturation

The tertiary structure of a globular protein is delicately held together by weak intramolecular attractions. Often, a modest change in temperature or pH will disrupt the tertiary structure and cause the protein to become **denatured**. Denaturation occurs under such mild conditions that covalent bonds aren't affected; the polypeptide primary structure remains intact, but the tertiary structure unfolds from a well-defined spherical shape to a randomly looped chain (Figure 27.13).

Denaturation is accompanied by changes in both physical and biological properties. Solubility is drastically decreased, as occurs when egg white is cooked and the albumins unfold and coagulate to an insoluble white mass. Most enzymes also lose all catalytic activity when denatured, since a precisely defined tertiary structure is required for their action.

Most, but not all, denaturation is irreversible. Eggs don't become uncooked when their temperature is lowered, and curdled milk doesn't become



**Figure 27.13** Schematic representation of protein denaturation. A globular protein loses its specific three-dimensional shape and becomes randomly looped.

homogeneous. Many cases have now been found, however, where spontaneous **renaturation** of an unfolded protein occurs. Renaturation is accompanied by a full recovery of biological activity in the case of enzymes, indicating that the protein has completely returned to its stable tertiary structure.

## 27.17 Summary and Key Words

**Proteins** are large biomolecules made up of  **$\alpha$ -amino acid residues** linked together by amide, or peptide, bonds. Twenty amino acids are commonly found in proteins; all are  $\alpha$ -amino acids and all except glycine have stereochemistry similar to that of L sugars.

Amino acids can be synthesized by several methods, including ammonolysis of  $\alpha$ -bromo acids, **reductive amination** of  $\alpha$ -keto acids, **Strecker reaction** of aldehydes with KCN/ $\text{NH}_4\text{Cl}$  followed by hydrolysis, and alkylation of diethyl **acetamidomalonnate**. Resolution of the synthetic racemate then provides the optically active amino acid.

Determining the structure of a large polypeptide or protein is carried out in several steps. The identity and amount of each amino acid present in a peptide can be determined by **amino acid analysis**. The peptide is first hydrolyzed to its constituent  $\alpha$ -amino acids, which are then chromatographically separated and identified. Next, the peptide is **sequenced**. **Edman degradation** by treatment with phenyl isothiocyanate cleaves off one residue from the N terminus of the peptide and forms an easily identifiable derivative of the N-terminal amino acid. A series of sequential Edman degradations allows the sequencing of a peptide chain up to 25 residues in length.

Peptide synthesis is made possible by the use of selective **protecting groups**. An N-protected amino acid having a free carboxyl group is coupled with an O-protected amino acid having a free amino group in the presence of dicyclohexylcarbodiimide (DCC). Amide formation occurs, the protecting groups are removed, and the sequence is repeated. Amines are usually

protected as their *tert*-butoxycarbonyl (BOC) derivatives, and acids are protected as esters. This synthetic sequence is often carried out by the **Merrifield solid-phase technique**, in which the peptide is esterified to an insoluble polymeric support.

**Proteins** are classified as either globular or fibrous, depending on their **secondary and tertiary structures**. **Fibrous proteins** such as  $\alpha$ -keratin are tough, rigid, and water-insoluble, and are used in nature for forming structures such as hair and nails. **Globular proteins** such as myoglobin are water-soluble, roughly spherical in shape, and are mobile within cells. Most of the 2000 or so known enzymes are globular proteins.

## 27.18 Summary of Reactions

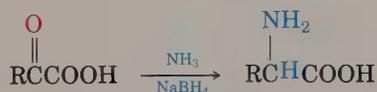
1. Amino acid synthesis
  - a. From  $\alpha$ -bromo acids (Section 27.4)



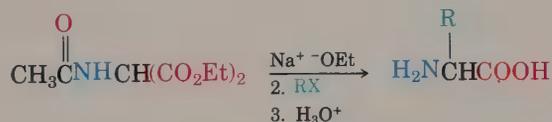
- b. Strecker synthesis (Section 27.4)



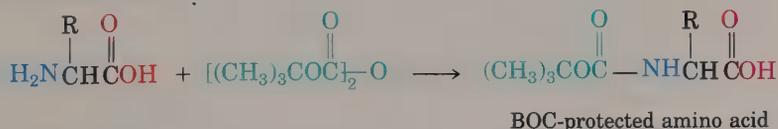
- c. Reductive amination (Sections 25.7 and 27.4)



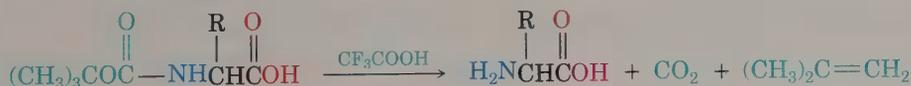
- d. Diethyl acetamidomalonate synthesis (Sections 22.9 and 27.4)



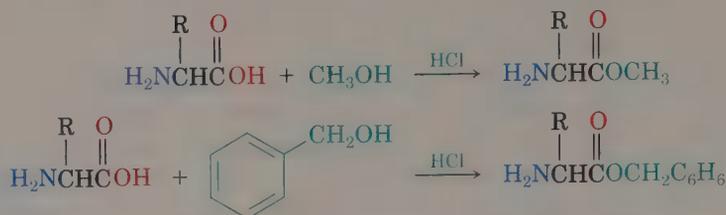
2. Peptide synthesis
  - a. Nitrogen protection (Section 27.11)



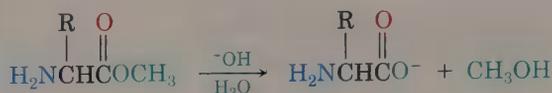
The BOC protecting group can be removed by acid treatment:



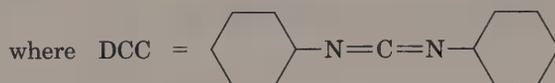
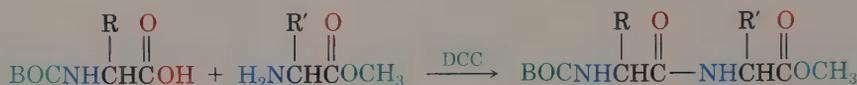
## b. Oxygen protection (Section 27.11)



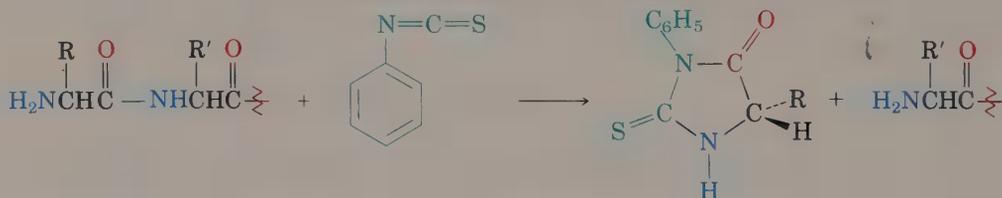
The ester protecting group can be removed by base hydrolysis:



## c. Amide bond formation (Section 27.11)

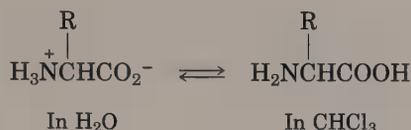


## 3. Peptide sequencing: Edman degradation (Section 27.9)

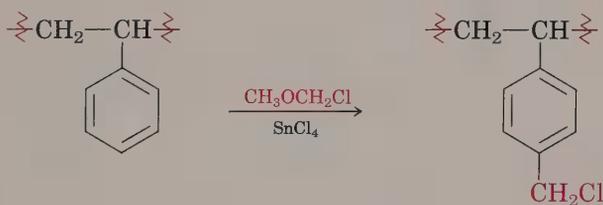


## ADDITIONAL PROBLEMS .....

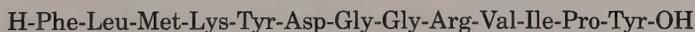
- 27.29** Only *S* amino acids occur in proteins, but several *R* amino acids are also found in nature. Thus, (*R*)-serine is found in earthworms and (*R*)-alanine is found in insect larvae. Draw Fischer projections of (*R*)-serine and (*R*)-alanine.
- 27.30** Draw a Fischer projection of (*S*)-proline.
- 27.31** Show the structures of these amino acids:
- |         |         |
|---------|---------|
| (a) Trp | (b) Ile |
| (c) Cys | (d) His |
- 27.32** Draw these amino acids in their zwitterionic forms:
- |              |               |
|--------------|---------------|
| (a) Tyrosine | (b) Threonine |
|--------------|---------------|
- 27.33** Explain the observation that amino acids exist as dipolar zwitterions in aqueous solution but exist largely as amino carboxylic acids in chloroform solution.



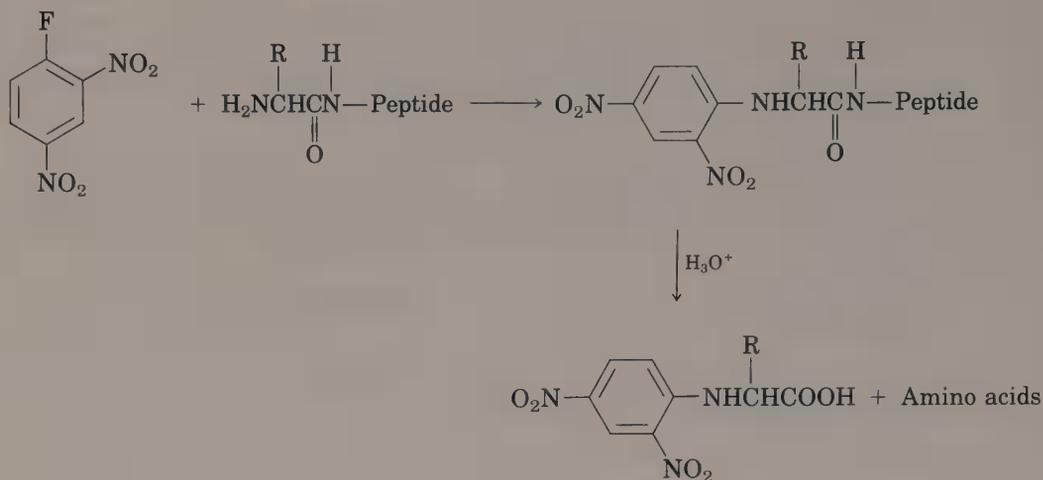
- 27.34** At what pH would you carry out an electrophoresis experiment if you wanted to separate a mixture of histidine, serine, and glutamic acid? Explain.
- 27.35** Define these terms:  
 (a) Amphoteric  
 (b) Isoelectric point  
 (c) Zwitterion
- 27.36** Using the three-letter code names for amino acids, write the structures of all possible peptides containing these amino acids:  
 (a) Val, Ser, Leu  
 (b) Ser, Leu<sub>2</sub>, Pro
- 27.37** Cytochrome *c* is an enzyme found in the cells of all aerobic organisms. Elemental analysis of cytochrome *c* shows that it contains 0.43% iron. What is the minimum molecular weight of this enzyme?
- 27.38** Predict the product of the reaction of valine with these reagents:  
 (a) CH<sub>3</sub>CH<sub>2</sub>OH, acid  
 (b) Di-*tert*-butyl dicarbonate  
 (c) KOH, H<sub>2</sub>O  
 (d) CH<sub>3</sub>COCl, pyridine; then H<sub>2</sub>O
- 27.39** Write full structures for these peptides:  
 (a) H-Val-Phe-Cys-Ala-OH  
 (b) H-Glu-Pro-Ile-Leu-OH
- 27.40** Show the steps involved in a synthesis of H-Phe-Ala-Val-OH using the Merrifield procedure.
- 27.41** Draw the structure of the phenylthiohydantoin product you would obtain by Edman degradation of these peptides:  
 (a) H-Ile-Leu-Pro-Phe-OH  
 (b) H-Asp-Thr-Ser-Gly-Ala-OH
- 27.42** The chloromethylated polystyrene resin used for Merrifield solid-phase peptide synthesis is prepared by treatment of polystyrene with chloromethyl methyl ether and a Lewis acid catalyst. Propose a mechanism for the reaction.



- 27.43** Which amide bonds in the following polypeptide are cleaved by trypsin? By chymotrypsin?



- 27.44** The Sanger end-group determination is sometimes used as an alternative to the Edman degradation. In the Sanger method, a peptide is allowed to react with 2,4-dinitrofluorobenzene, the peptide is hydrolyzed, and the N-terminal amino acid is identified by separation as its *N*-2,4-dinitrophenyl derivative. Propose a mechanism to account for the initial reaction between peptide and dinitrofluorobenzene.

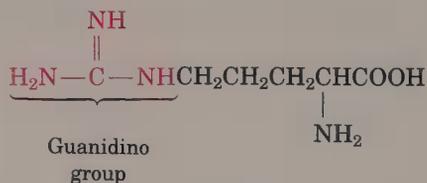


- 27.45 Would you foresee any problems in using the Sanger end-group determination method (Problem 27.44) on a peptide such as H-Gly-Pro-Lys-Ile-OH? Explain.
- 27.46 When  $\alpha$ -amino acids are treated with dicyclohexylcarbodiimide (DCC), 2,5-diketopiperazines result. Propose a mechanism for this reaction.



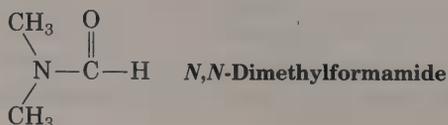
A 2,5-diketopiperazine

- 27.47 Arginine, which contains a guanidino functional group in its side chain, is by far the most basic of the 20 common amino acids. How can you account for this basicity? Use resonance structures to see how the protonated guanidino group is stabilized.



Arginine

- 27.48 Evidence for restricted rotation around amide CO-N bonds comes from NMR studies. At room temperature, the  $^1\text{H}$  NMR spectrum of *N,N*-dimethylformamide shows three peaks: 2.9  $\delta$  (singlet, 3 H), 3.0  $\delta$  (singlet, 3 H), 8.0  $\delta$  (singlet, 1 H). As the temperature is raised, however, the two singlets at 2.9  $\delta$  and 3.0  $\delta$  slowly merge. At 180°C, the  $^1\text{H}$  NMR spectrum shows only two peaks: 2.95  $\delta$  (singlet, 6 H) and 8.0  $\delta$  (singlet, 1 H). Explain this temperature-dependent behavior.

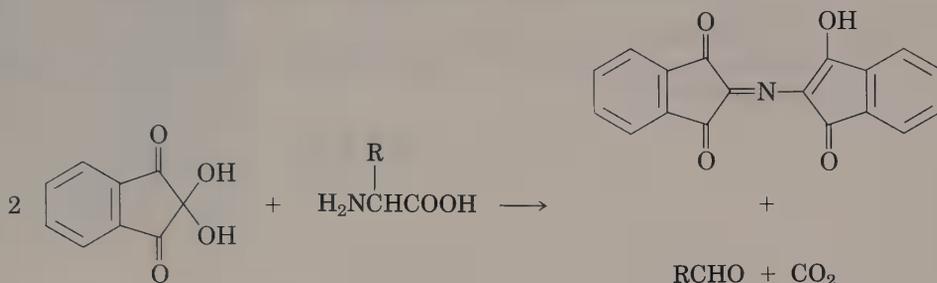


- 27.49 An octapeptide shows the composition Asp, Gly<sub>2</sub>, Leu, Phe, Pro<sub>2</sub>, Val on amino acid analysis. Edman analysis shows a glycine N-terminal group, and carboxypeptidase cleavage produces leucine as the first amino acid to appear. Acidic hydrolysis gives the following fragments:

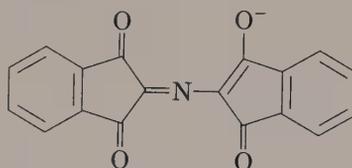
1. H-Val-Pro-Leu-OH
2. H-Gly-OH
3. H-Gly-Asp-Phe-Pro-OH
4. H-Phe-Pro-Val-OH

Propose a structure for the starting octapeptide.

- 27.50 Propose a mechanism to account for the reaction of ninhydrin with an  $\alpha$ -amino acid:



- 27.51 Draw as many resonance forms as you can for the purple anion obtained by reaction of ninhydrin with an  $\alpha$ -amino acid (Problem 27.50).



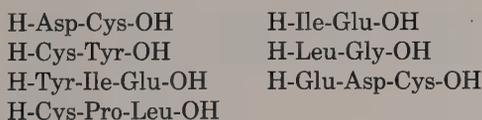
- 27.52 Look up the structure of human insulin (Figure 27.9), and indicate where in each chain the molecule is cleaved by trypsin and chymotrypsin.
- 27.53 What is the structure of a nonapeptide that gives the following fragments when cleaved?

Trypsin cleavage: H-Val-Val-Pro-Tyr-Leu-Arg-OH and H-Ser-Ile-Arg-OH

Chymotrypsin cleavage: H-Leu-Arg-OH and H-Ser-Ile-Arg-Val-Val-Pro-Tyr-OH

- 27.54 Oxytocin, a nonapeptide hormone secreted by the pituitary gland, functions by stimulating uterine contraction and lactation during childbirth. Its sequence was determined from the following evidence:

1. Oxytocin is a cyclic compound containing a disulfide bridge between two cysteine residues.
2. When the disulfide bridge is reduced, oxytocin has the constitution Asn, Cys<sub>2</sub>, Gln, Gly, Ile, Leu, Pro, Tyr.
3. Partial hydrolysis of reduced oxytocin yields seven fragments:

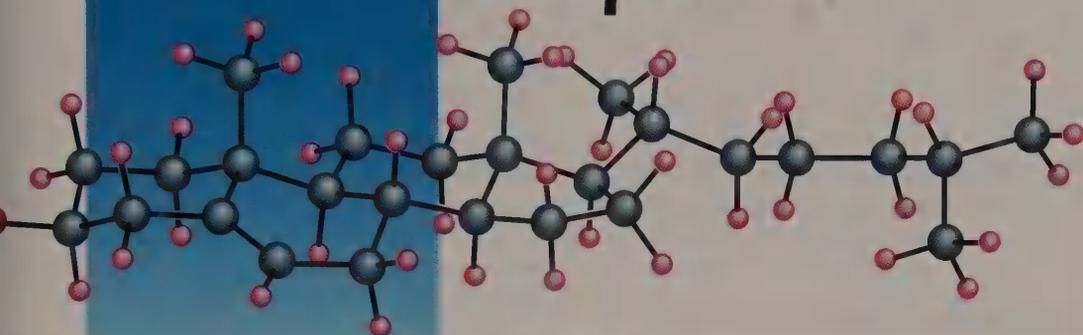


4. Gly is the C-terminal group.
5. Both Glu and Asp are present as their side-chain amides (Gln and Asn) rather than as free side-chain acids.

On the basis of this evidence, what is the amino acid sequence of reduced oxytocin? What is the structure of oxytocin itself?

- 27.55** *Aspartame*, a nonnutritive sweetener marketed under the trade name Nutra-Sweet, is the methyl ester of a simple dipeptide, H-Asp-Phe-OCH<sub>3</sub>.
- (a) Draw the full structure of aspartame.
  - (b) The isoelectric point of aspartame is 5.9. Draw the principal structure present in aqueous solution at this pH.
  - (c) Draw the principal form of aspartame present at physiological pH = 7.6.
- 27.56** Refer to Figure 27.6 and propose a mechanism for the final step in the Edman degradation—the acid-catalyzed rearrangement of the thiazolinone to the *N*-phenylthiohydantoin.

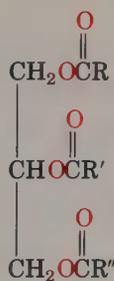
# Chapter 28



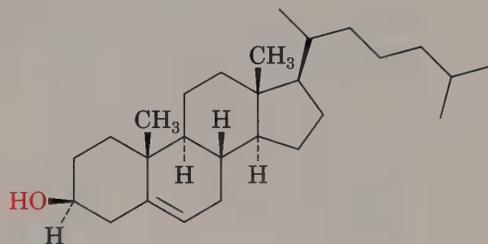
## Lipids

**Lipids** are naturally occurring organic molecules isolated from cells and tissues by extraction with nonpolar organic solvents. Since they usually have large hydrocarbon portions in their structures, lipids are insoluble in water but soluble in organic solvents. Note that this definition differs from the sort used for carbohydrates and proteins. Lipids are defined by physical property (solubility) rather than by structure.

Lipids are classified into two general types: those like fats and waxes that contain hydrolyzable ester linkages, and those like cholesterol and other steroids that don't have ester linkages and can't be hydrolyzed.



Animal fat, an ester  
(R, R', R'' = C<sub>11</sub>-C<sub>19</sub> chains)

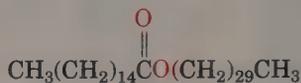


Cholesterol

### 28.1 Waxes, Fats, and Oils

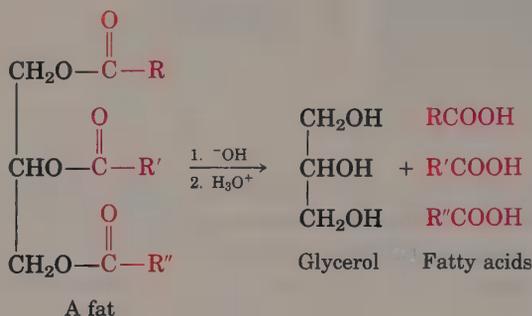
Waxes are mixtures of esters between long-chain carboxylic acids and long-chain alcohols. The carboxylic acid usually has an even number of carbons from 16 through 36, while the alcohol has an even number of carbons from 24 through 36. One of the major components of beeswax, for instance, is

triacontyl hexadecanoate, the ester from the  $C_{30}$  alcohol triacontanol and the  $C_{16}$  acid hexadecanoic acid. The waxy protective coatings on most fruits, berries, leaves, and animal furs have similar structures.



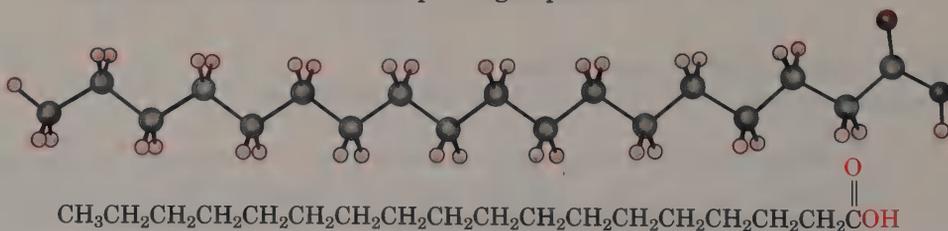
Triacontyl hexadecanoate (from beeswax)

Animal fats and vegetable oils are the most widely occurring lipids. Although they appear different—animal fats like butter and lard are solids, whereas vegetable oils like corn and peanut oil are liquid—their structures are closely related. Chemically, fats and oils are **triacylglycerols**, also called *triglycerides*—that is, triesters of glycerol with three long-chain carboxylic acids. Thus, hydrolysis of a fat or oil with aqueous sodium hydroxide yields glycerol and three **fatty acids**:



The fatty acids obtained by hydrolysis of triacylglycerols are generally unbranched and contain an even number of carbon atoms between 12 and 20. If double bonds are present, they usually have *Z* (cis) geometry. The three fatty acids of a specific triacylglycerol molecule aren't usually the same, and the fat or oil from a given source is likely to be a complex mixture of many different triacylglycerols. Table 28.1 lists some of the commonly occurring fatty acids, and Table 28.2 lists the approximate composition of some fats and oils from different sources.

About 40 different fatty acids occur naturally. Palmitic acid ( $C_{16}$ ) and stearic acid ( $C_{18}$ ) are the most abundant saturated acids; oleic and linoleic acids (both  $C_{18}$ ) are the most abundant unsaturated ones. Oleic acid is monounsaturated since it has only one double bond, whereas linoleic, linolenic, and arachidonic acid are polyunsaturated fatty acids, or PUFAs, which have more than one double bond. Linoleic and linolenic acids are essential in the human diet; infants grow poorly and develop skin lesions if fed a diet of nonfat milk for prolonged periods.



Palmitic acid

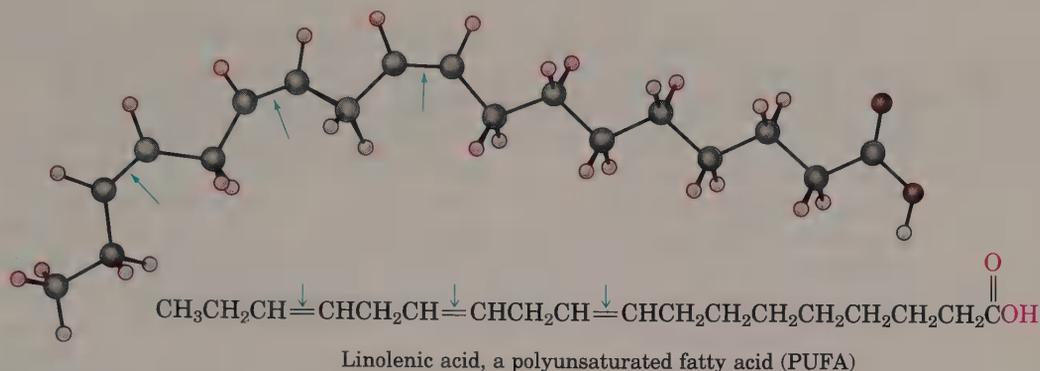


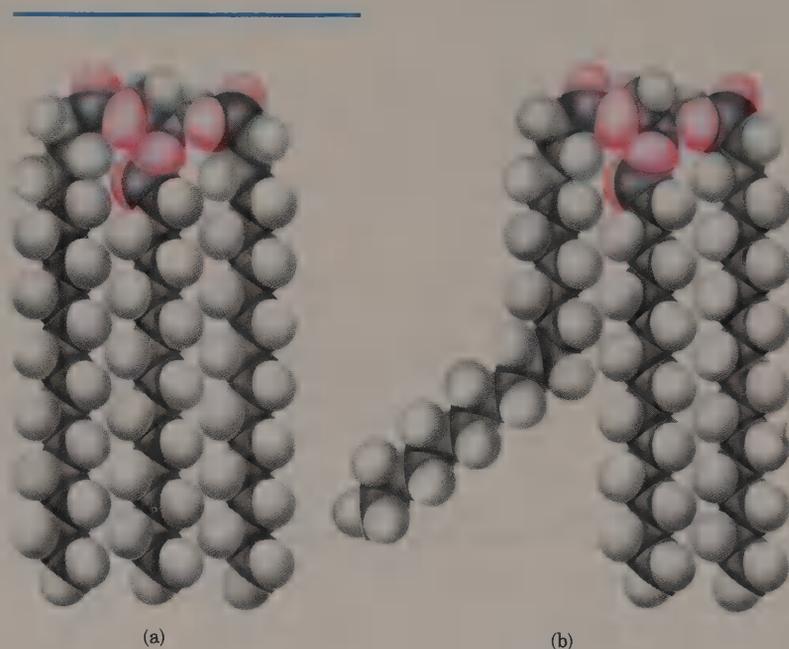
TABLE 28.1 Structures of Some Common Fatty Acids

Name	Carbons	Structure	Melting point (°C)
<b>Saturated</b>			
Lauric	12	$\text{CH}_3(\text{CH}_2)_{10}\text{COOH}$	44
Myristic	14	$\text{CH}_3(\text{CH}_2)_{12}\text{COOH}$	58
Palmitic	16	$\text{CH}_3(\text{CH}_2)_{14}\text{COOH}$	63
Stearic	18	$\text{CH}_3(\text{CH}_2)_{16}\text{COOH}$	70
Arachidic	20	$\text{CH}_3(\text{CH}_2)_{18}\text{COOH}$	75
<b>Unsaturated</b>			
Palmitoleic	16	$\text{CH}_3(\text{CH}_2)_5\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ (cis)	32
Oleic	18	$\text{CH}_3(\text{CH}_2)_7\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ (cis)	4
Ricinoleic	18	$\text{CH}_3(\text{CH}_2)_5\text{CH}(\text{OH})\text{CH}_2\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ (cis)	5
Linoleic	18	$\text{CH}_3(\text{CH}_2)_4\text{CH}=\text{CHCH}_2\text{CH}=\text{CH}(\text{CH}_2)_7\text{COOH}$ (cis,cis)	-5
Arachidonic	20	$\text{CH}_3(\text{CH}_2)_4\text{CH}=\text{CHCH}_2)_4\text{CH}_2\text{CH}_2\text{COOH}$ (all cis)	-50

TABLE 28.2 Approximate Fatty Acid Composition of Some Fats and Oils

Source	Saturated fatty acids (%)				Unsaturated fatty acids (%)		
	C <sub>12</sub> Lauric	C <sub>14</sub> Myristic	C <sub>16</sub> Palmitic	C <sub>18</sub> Stearic	C <sub>18</sub> Oleic	C <sub>18</sub> Ricinoleic	C <sub>18</sub> Linoleic
<b>Animal fat</b>							
Lard	—	1	25	15	50	—	6
Butter	2	10	25	10	25	—	5
Human fat	1	3	25	8	46	—	10
Whale blubber	—	8	12	3	35	—	10
<b>Vegetable oil</b>							
Coconut	50	18	8	2	6	—	1
Corn	—	1	10	4	35	—	45
Olive	—	1	5	5	80	—	7
Peanut	—	—	7	5	60	—	20
Linseed	—	—	5	3	20	—	20
Castor bean	—	—	—	1	8	85	4

The data in Table 28.1 show that unsaturated fatty acids generally have lower melting points than their saturated counterparts have, a trend that also holds true for triacylglycerols. Since vegetable oils generally have a higher proportion of unsaturated to saturated fatty acids than animal fats have (Table 28.2), they have lower melting points. The difference in melting points between fats and oils is a consequence of the difference in their structures. Saturated fats have a uniform shape that allows them to pack together easily in a crystal lattice. Carbon-carbon double bonds in unsaturated vegetable oils, however, introduce bends and kinks into the hydrocarbon chains, making crystal formation difficult. The more double bonds there are, the harder it is for the molecules to crystallize and the lower the melting point of the oil. Figure 28.1 illustrates this effect with space-filling molecular models.



**Figure 28.1** Space-filling molecular models of (a) saturated and (b) unsaturated triacylglycerols. There is an unsaturated fatty acid in (b) that prevents the molecule from adopting a regular shape and crystallizing easily.

The carbon-carbon double bonds in vegetable oils can be reduced by catalytic hydrogenation (Section 7.7) to produce saturated solid or semisolid fats. Margarine and solid cooking fats such as Crisco are produced by hydrogenating soybean, peanut, or cottonseed oil until the proper consistency is obtained.

## PROBLEM .....

- 28.1 The carnauba wax used in floor and furniture polishes contains, among other things, an ester of a  $C_{32}$  straight-chain alcohol with a  $C_{20}$  straight-chain carboxylic acid. Draw its structure.

## PROBLEM .....

- 28.2 Draw structures of these molecules. Which would you expect to be higher melting?  
 (a) Glyceryl tripalmitate (b) Glyceryl trioleate

## PROBLEM .....

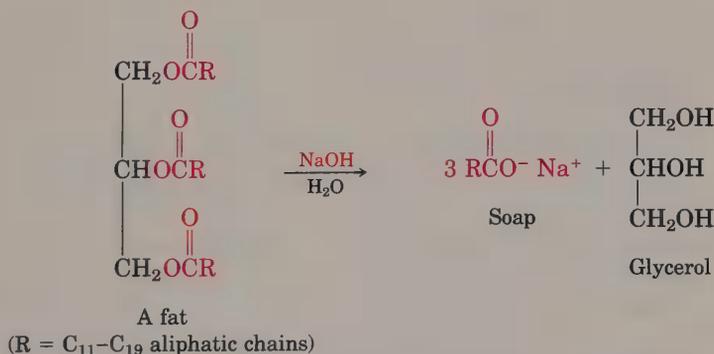
- 28.3 Stearolic acid,  $C_{18}H_{32}O_2$ , yields stearic acid on catalytic hydrogenation and undergoes oxidative cleavage with ozone to yield nonanoic acid and nonanedioic acid. What is the structure of stearolic acid?
- .....

## 28.2 Soaps

---

Soap has been known since at least 600 B.C., when the Phoenicians reportedly prepared a curdy material by boiling goat fat with extracts of wood ash. The cleansing properties of soap weren't generally recognized, however, and the use of soap did not become widespread until the eighteenth century.

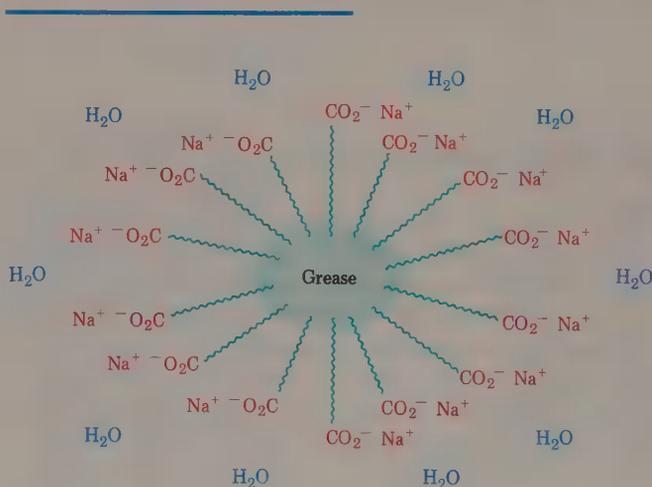
Chemically, **soap** is a mixture of the sodium or potassium salts of long-chain fatty acids produced by hydrolysis (*saponification*) of animal fat with alkali. Wood ash was used as a source of alkali until the mid-1800s, when NaOH became commercially available.



Crude soap curds contain glycerol and excess alkali as well as soap, but purification can be effected by boiling with a large amount of water and adding NaCl to precipitate the pure sodium carboxylate salts. The smooth soap that precipitates is dried, perfumed, and pressed into bars for household use. Dyes are added to make colored soaps, antiseptics are added for medicated soaps, pumice is added for scouring soaps, and air is blown in for soaps that float. Regardless of these extra treatments and regardless of price, though, all soaps are basically the same.

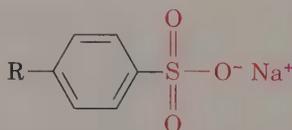
Soaps act as cleansers because the two ends of a soap molecule are so different. The sodium salt end of the long-chain molecule is ionic and therefore *hydrophilic* (water-loving); it tries to dissolve in water. The long hydrocarbon chain portion of the molecule, however, is nonpolar and *lipophilic* (fat-loving); it tries to dissolve in grease. The net effect of these two opposing tendencies is that soaps are attracted to both grease and water and are therefore valuable as cleansers.

When soaps are dispersed in water, the long hydrocarbon tails cluster together in a lipophilic ball, while the ionic heads on the surface of the cluster stick out into the water layer. These spherical clusters, called **micelles**, are shown schematically in Figure 28.2. Grease and oil droplets are solubilized in water when they are coated by the nonpolar tails of soap molecules in the center of micelles. Once solubilized, the grease and dirt can be rinsed away.



**Figure 28.2** A soap micelle solubilizing a grease particle in water.

Soaps make life more pleasant than it might otherwise be, but they have certain drawbacks. In hard water, which contains metal ions, soluble sodium carboxylates are converted into insoluble magnesium and calcium salts, leaving the familiar ring of scum around bathtubs and the gray tinge on white clothes. Chemists have circumvented these problems by synthesizing a class of synthetic detergents based on salts of long-chain alkylbenzenesulfonic acids. The principle of synthetic detergents is identical to the principle of soaps: The alkylbenzene end of the molecule is lipophilic and



A synthetic detergent  
(R = a mixture of  $C_{12}$  aliphatic chains)

attracts grease, but the sulfonate salt end is ionic and is attracted to water. Unlike soaps, though, sulfonate detergents don't form insoluble metal salts in hard water and don't leave an unpleasant scum.

PROBLEM .....

28.4 Draw the structure of magnesium oleate, a component of bathtub scum.

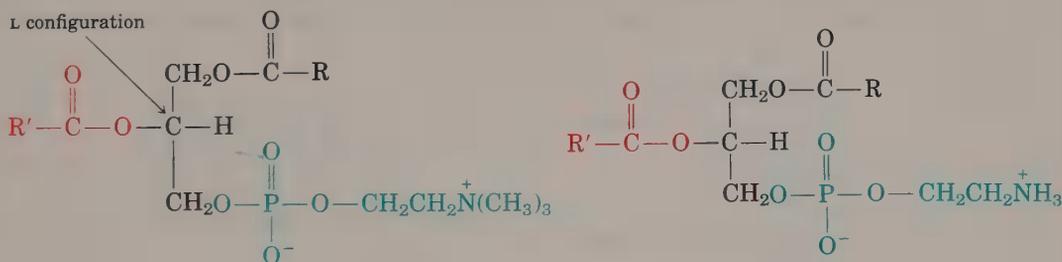
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## 28.3 Phospholipids

Just as waxes, fats, and oils are esters of carboxylic acids, **phospholipids** are esters of phosphoric acid,  $\text{H}_3\text{PO}_4$ . There are two main kinds of phospholipids: *phosphoglycerides* and *sphingolipids*.

**Phosphoglycerides** are closely related to fats and oils in that they contain a glycerol backbone linked by ester bonds to two fatty acids and one phosphoric acid. Although the fatty acid residues can be any of the  $\text{C}_{12}$ – $\text{C}_{20}$  units normally present in fats, the acyl group at C1 is usually saturated, and that at C2 is usually unsaturated. The phosphate group at C3 is also bound by a separate ester link to an amino alcohol such as choline,  $\text{HOCH}_2\text{CH}_2\text{N}^+(\text{CH}_3)_3$ , or ethanolamine,  $\text{HOCH}_2\text{CH}_2\text{NH}_2$ .

The most important phosphoglycerides are the *lecithins* and the *cephalins*. Note that these compounds are chiral and that they have the L, or R, configuration at C2.

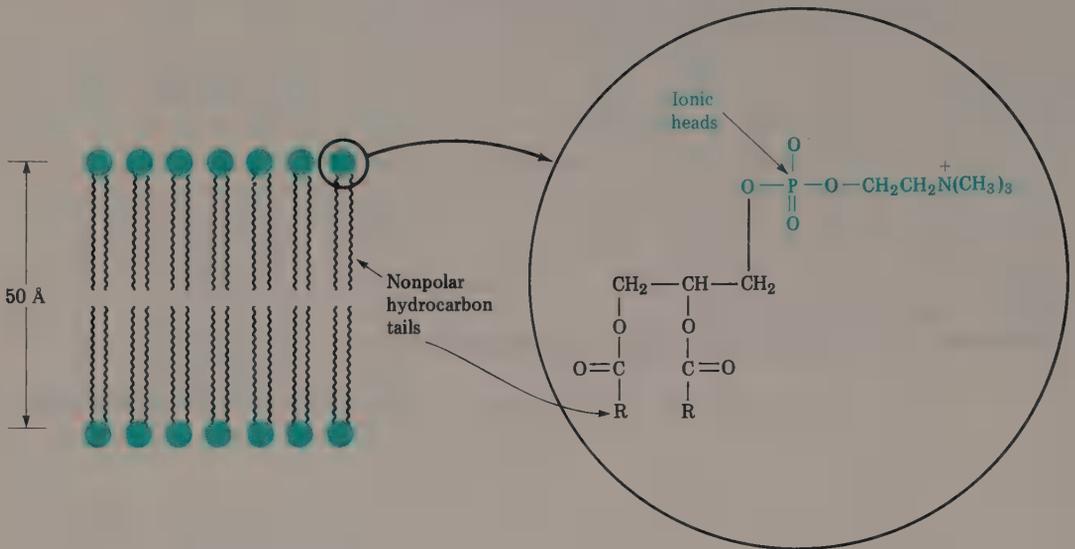


Phosphatidylcholine, a lecithin

Phosphatidylethanolamine, a cephalin

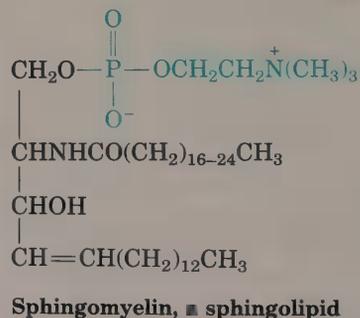
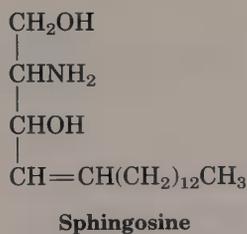
where R is saturated and R' is unsaturated

Found widely in both plant and animal tissues, phosphoglycerides are the major lipid component of cell membranes (approximately 40%). Like soaps, phosphoglycerides have a long, nonpolar hydrocarbon tail bound to a polar ionic head (the phosphate group). Cell membranes are composed in large part of phosphoglycerides oriented into a **lipid bilayer** about 50 Å thick. As shown in Figure 28.3, the lipophilic tails aggregate in the center of the bilayer in much the same way that soap tails aggregate in the center of a micelle. This bilayer serves as an effective barrier to the passage of water, ions, and other components into and out of the cell.



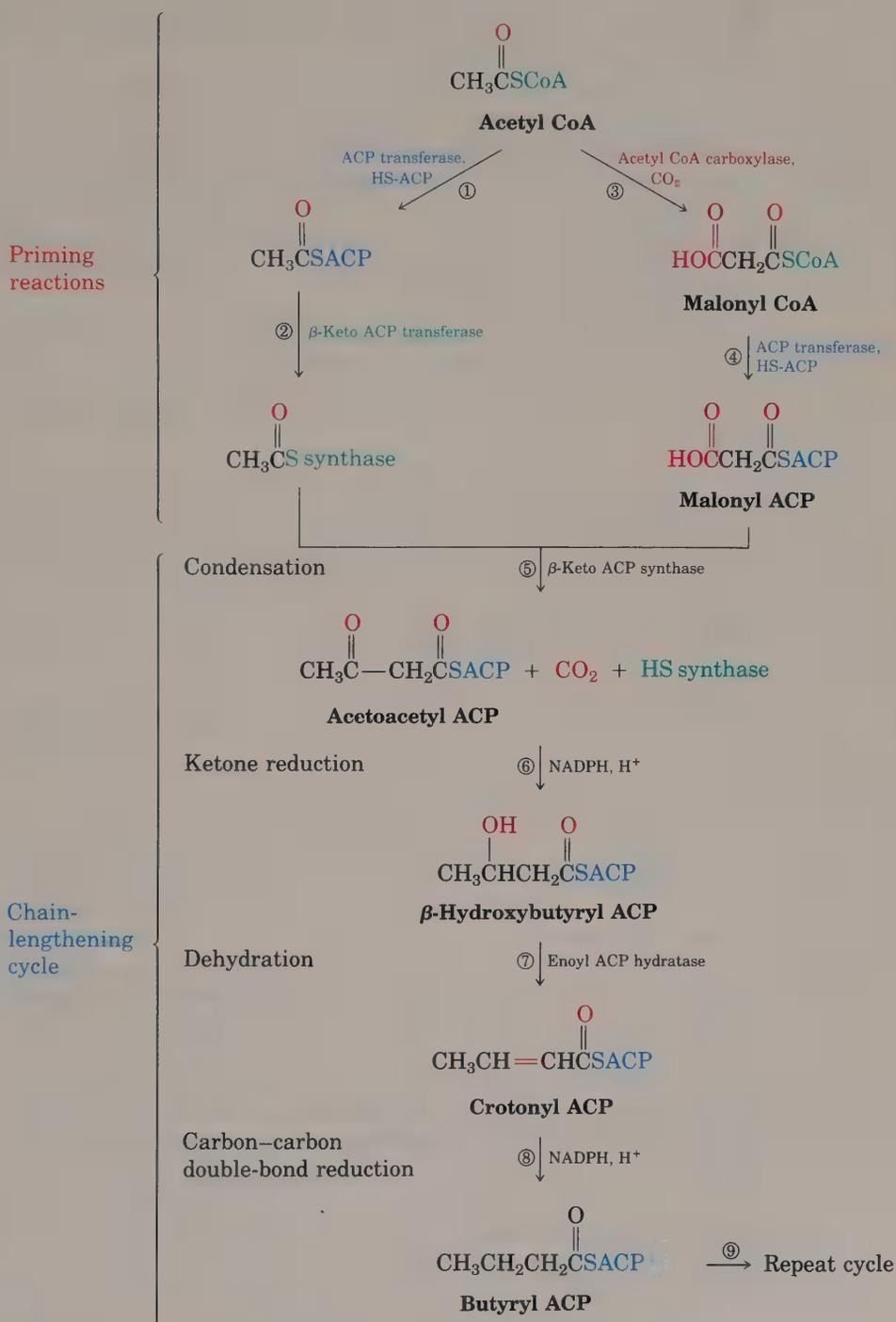
**Figure 28.3** Aggregation of phosphoglycerides into the lipid bilayer that comprises cell membranes.

The second major group of phospholipids are the **sphingolipids**. These complex lipids, which have *sphingosine* or a related dihydroxyamine as their backbones, are constituents of plant and animal cell membranes. They're particularly abundant in brain and nerve tissue, where compounds called *sphingomyelins* are a major constituent of the coating around nerve fibers.



## 28.4 Biosynthesis of Fatty Acids

One of the most striking features of the fatty acids shown in Table 28.1 is that all have an even number of carbon atoms. The reason is that all are derived biosynthetically from the simple two-carbon precursor, acetic acid. The pathway by which organisms synthesize fat (*lipogenesis*) is shown in Figure 28.4.



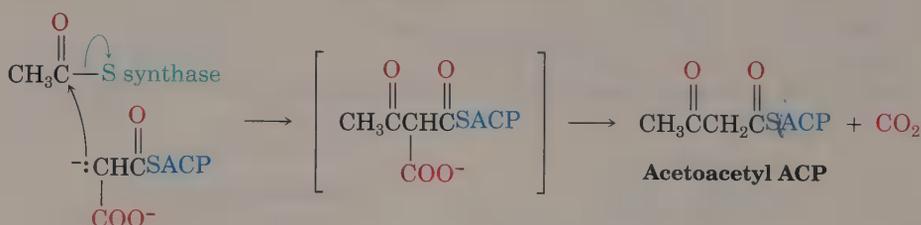
**Figure 28.4** Biological pathway for fatty acid synthesis.

The starting material for fatty acid synthesis is the thiol ester acetyl CoA (Section 23.15), prepared in nature from acetic acid. The fatty acid synthetic pathway begins with several *priming reactions* that convert acetyl CoA into more reactive species.

The first two steps in the priming sequence convert acetyl CoA into *acetyl synthase*. This reactive acylating agent is capable of transferring the acetyl group to a nucleophile by nucleophilic acyl substitution reaction (Section 21.2). In step 1, a reaction catalyzed by the enzyme *acyl carrier protein (ACP) transferase* exchanges the thiol ester linkage of acetyl CoA for a different and somewhat more reactive thiol ester bond to ACP. Step 2 involves a further exchange of thiol ester linkages, resulting in the formation of acetyl synthase.

The next two steps in the priming sequence again start with acetyl CoA. In step 3, acetyl CoA is carboxylated by reaction with  $\text{CO}_2$  and the enzyme acetyl CoA carboxylase to yield *malonyl CoA*. Step 4 is another thiol ester exchange reaction that converts malonyl CoA into the more reactive *malonyl ACP*.

The key carbon–carbon bond-forming reaction used to build the fatty acid chain occurs in step 5. This key step is simply a Claisen condensation (Section 23.9) between acetyl synthase as the electrophilic acceptor component and malonyl ACP as the nucleophilic donor component. An enolate ion derived from the doubly activated  $-\text{CH}_2-$  group of malonyl ACP adds to the carbonyl group of acetyl synthase, yielding an intermediate  $\beta$ -keto acid that loses carbon dioxide to give the four-carbon product, *acetoacetyl ACP*.



Step 6 involves the NADPH (nicotinamide adenine dinucleotide phosphate) enzyme-catalyzed reduction of the ketone carbonyl group in acetoacetyl ACP to an alcohol,  *$\beta$ -hydroxybutyryl ACP*, which is dehydrated in step 7 to yield *crotonyl ACP*. The carbon–carbon double bond of crotonyl ACP is further reduced by NADPH in step 8 to yield *butyryl ACP*. The net effect of these eight separate steps is to take two acetic acid units and combine them into a single four-carbon butanoic acid unit. Further repetition of the cycle by condensation of butyryl synthase with another malonyl ACP yields a six-carbon unit, and still further repetitions add two more carbon atoms to the chain each time.

Fatty acid biosynthesis is a lengthy process—some 64 separate steps are required for a cell to produce stearic acid—but all the individual steps are simple transformations of the same sort a chemist might do in the laboratory. We've seen them all before: nucleophilic acyl substitution, Claisen condensation, ketone reduction, dehydration, alkene hydrogenation. Clearly, biological chemistry and laboratory organic chemistry follow the same rules.

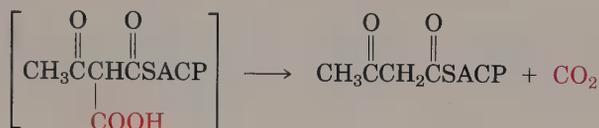
## PROBLEM.....

- 28.5 Show a likely reaction mechanism for the transformation of acetyl CoA into acetyl ACP.



## PROBLEM.....

- 28.6 Show a likely mechanism for the ready decarboxylation of the acylmalonyl ACP intermediate formed in step 5 of Figure 28.4.



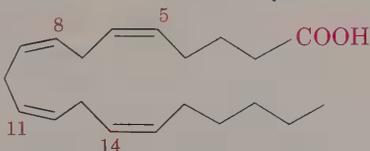
## PROBLEM.....

- 28.7 Evidence for the role of acetate in fatty acid biosynthesis comes from isotope-labeling experiments. If acetate labeled with  $^{14}\text{C}$  in the methyl group ( $^{14}\text{CH}_3\text{COOH}$ ) were incorporated into fatty acids, at what positions in the fatty acid chain would you expect the  $^{14}\text{C}$  label to appear?

## 28.5 Prostaglandins

Few compounds have caused as much excitement among medical researchers in the past decade as the prostaglandins. First isolated by Sune Bergstrom,<sup>1</sup> Bengt Samuelsson,<sup>2</sup> and their collaborators at the Karolinska Institute in Sweden, these simple lipids are synthesized in nature from the  $\text{C}_{20}$  fatty acid, arachidonic acid. The name *prostaglandin* derives from the fact that these compounds were first thought to be produced by the prostate gland, but they have subsequently been shown to be present in small amounts in all body tissues and fluids.

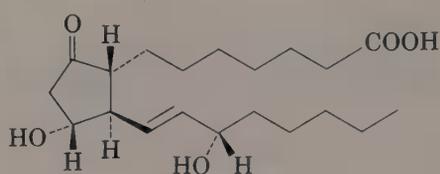
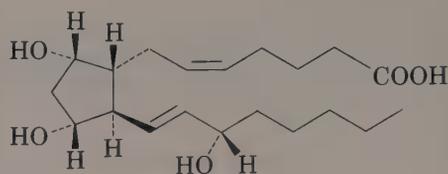
The **prostaglandins** are simple in structure. All are  $\text{C}_{20}$  carboxylic acids that contain a five-membered ring with two long side chains. They differ only in the number of oxygen atoms and in the number of double bonds present. Prostaglandin  $\text{E}_1$  ( $\text{PGE}_1$ ) and prostaglandin  $\text{F}_{2\alpha}$  ( $\text{PGF}_{2\alpha}$ ) are representative structures:



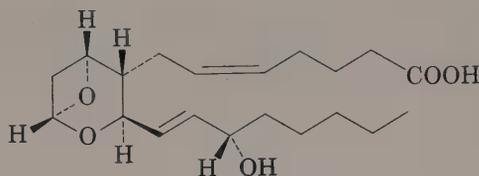
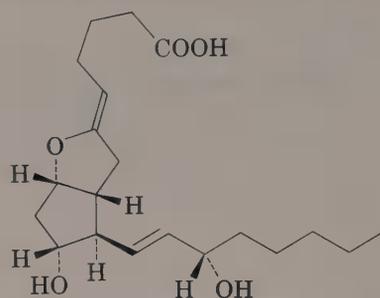
Arachidonic acid,  
(5*Z*,8*Z*,11*Z*,14*Z*)-eicosatetraenoic acid

<sup>1</sup>Sune Bergstrom (1916– ); M.D. Karolinska Institute; professor, Karolinska Institute; Nobel Prize (1982).

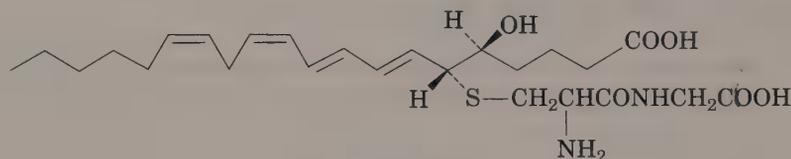
<sup>2</sup>Bengt Samuelsson (1934– ); M.D. Lund; professor, Karolinska Institute; Nobel Prize (1982).

Prostaglandin E<sub>1</sub>Prostaglandin F<sub>2α</sub>

The several dozen known prostaglandins have an extraordinarily wide range of biological effects. They can lower blood pressure, affect blood-platelet aggregation during clotting, lower gastric secretions, control inflammation, affect kidney function, affect reproductive systems, and stimulate uterine contractions during childbirth. In addition, compounds that are closely related to the prostaglandins have still other effects. Interest has centered particularly on the thromboxanes, on prostacyclin, and on the leukotrienes, whose release in the body appears to trigger the asthmatic response.

Thromboxane A<sub>2</sub>

Prostacyclin

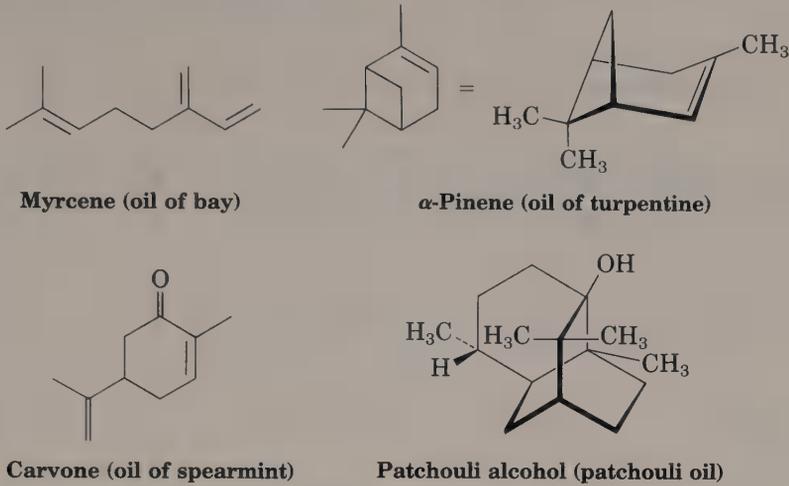
Leukotriene D<sub>4</sub>

Prostaglandins are being actively exploited in medicine, where their remarkable biological properties will surely lead to valuable new drugs.

## 28.6 Terpenes

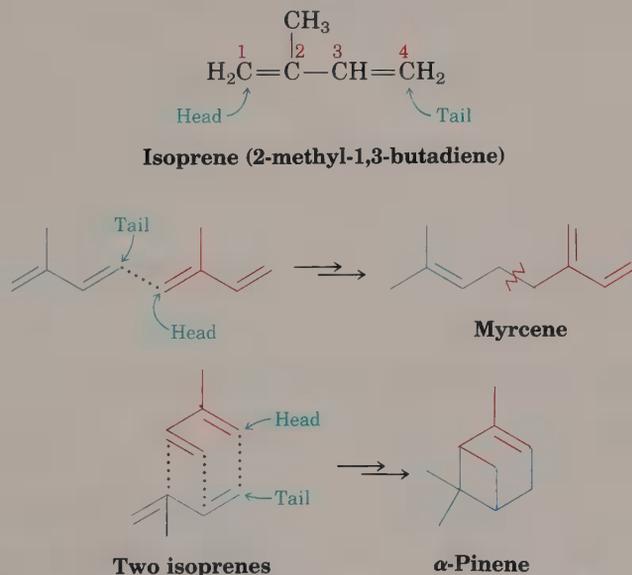
It has been known for centuries that codistillation of many plant materials with steam (called **steam distillation**) produces a fragrant mixture of liquids known as plant **essential oils**. For thousands of years, these plant extracts have been used as medicines, spices, and perfumes. The investigation of essential oils also played a major role in the emergence of organic chemistry as a science during the nineteenth century.

Chemically, the essential oils of plants consist largely of mixtures of lipids called **terpenes**. Terpenes are small organic molecules that have an immense diversity of structure. Thousands of different terpenes are known; some are hydrocarbons, and others contain oxygen; some are open-chain molecules, and others contain rings. Figure 28.5 gives some examples.



**Figure 28.5** The structures of some terpenes isolated from essential oils.

All terpenes are related, regardless of their apparent structural differences. According to the **isoprene rule** proposed by Ruzicka,<sup>3</sup> terpenes can be considered to arise from head-to-tail joining of simple five-carbon isoprene (2-methyl-1,3-butadiene) units. Carbon 1 is called the head of the isoprene unit, and carbon 4 is called the tail. For example, myrcene contains two isoprene units joined head to tail, forming an eight-carbon chain with two one-carbon branches.  $\alpha$ -Pinene similarly contains two isoprene units assembled into a more complex cyclic structure.



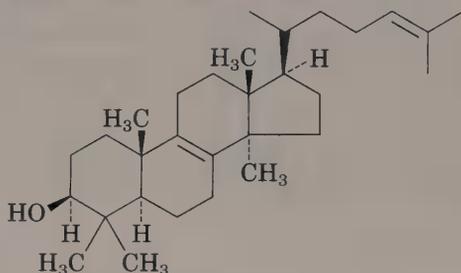
<sup>3</sup>Leopold Ruzicka (1887–1976); b. Vukovar (now Yugoslavia); Ph.D., 1910, Karlsruhe (Staudinger); professor, Swiss Federal Institute (E.T.H.) Zürich (1923–1926, 1929–1957); Nobel Prize (1939).

Terpenes are classified according to the number of 5-carbon isoprene units they contain. Thus, **monoterpenes** are 10-carbon substances biosynthesized from two isoprene units, **sesquiterpenes** are 15-carbon molecules from three isoprene units, and so on (Table 28.3).

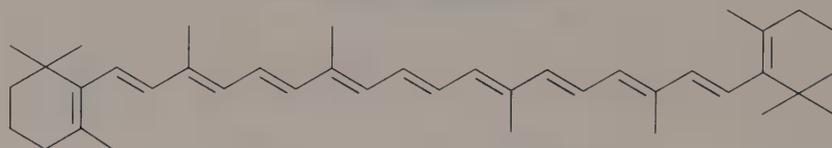
**TABLE 28.3** Classification of Terpenes

Carbon atoms	Isoprene units	Classification
10	2	Monoterpenes
15	3	Sesquiterpenes
20	4	Diterpenes
25	5	Sesterterpenes
30	6	Triterpenes
40	8	Tetraterpenes

Mono- and sesquiterpenes are found primarily in plants, but the higher terpenes occur in both plants and animals. Many of the higher terpenes have important biological roles. For example, the triterpene *lanosterol* is the precursor from which all steroid hormones are made in nature, and the tetraterpene  *$\beta$ -carotene* is a major dietary source of vitamin A.



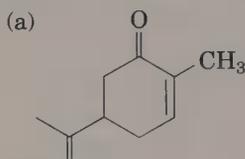
Lanosterol, a triterpene (C<sub>30</sub>)



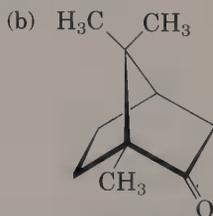
$\beta$ -Carotene, a tetraterpene (C<sub>40</sub>)

**PROBLEM** .....

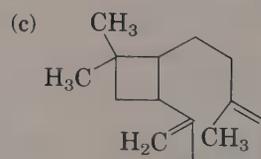
**28.8** Show the positions of the isoprene units in these terpenes:



Carvone (spearmint oil)



Camphor

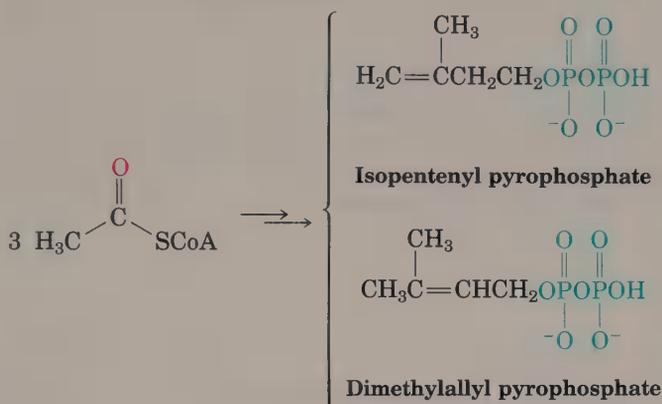


Caryophyllene (cloves)

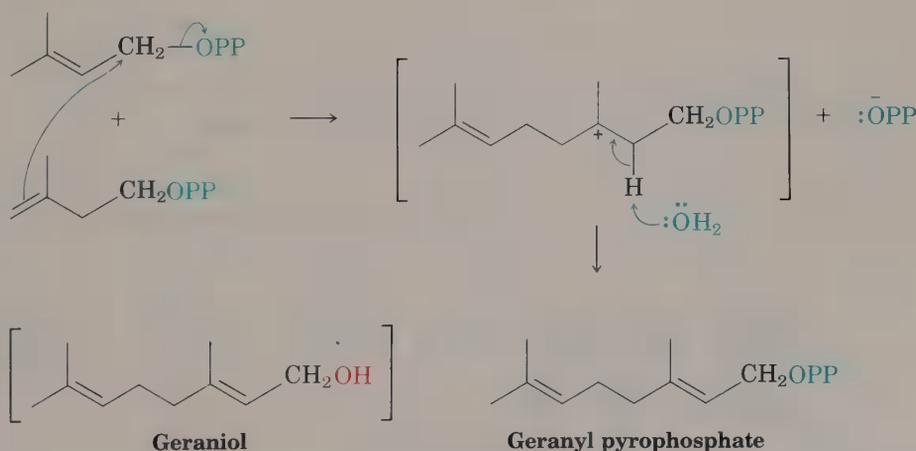
.....

## 28.7 Biosynthesis of Terpenes

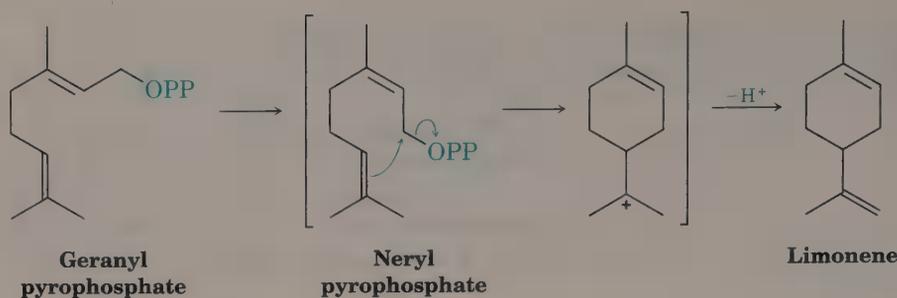
The isoprene rule is a convenient formalism, but isoprene itself is not the biological precursor of terpenes. Nature instead uses two isoprene “equivalents,” isopentenyl pyrophosphate and dimethylallyl pyrophosphate, for the biosynthesis of terpenes. These five-carbon molecules are themselves made from condensation of three acetyl CoA units.



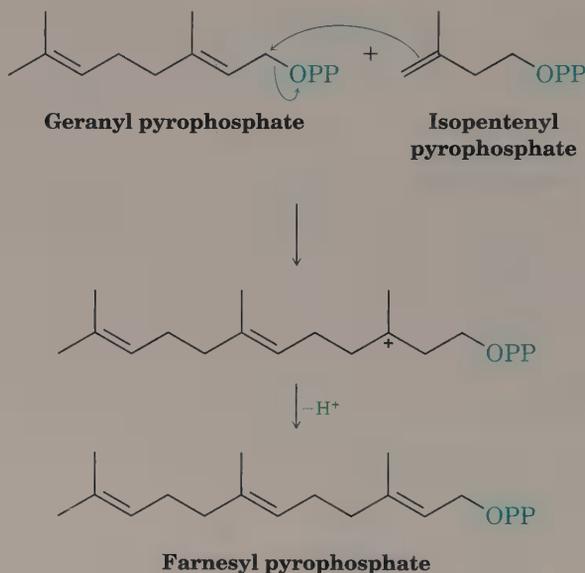
Dimethylallyl pyrophosphate is an excellent alkylating agent since the primary, allylic pyrophosphate group (abbreviated OPP) is a good leaving group in nucleophilic substitution reactions. We can therefore imagine a displacement of this leaving group by the nucleophilic double bond of isopentenyl pyrophosphate. Loss of a proton from the carbocationic reaction intermediate then leads to the head-to-tail coupled 10-carbon unit, *geranyl pyrophosphate*. The corresponding alcohol, geraniol, is itself a fragrant terpene that occurs in rose oil.



Geranyl pyrophosphate is the precursor of all other monoterpenes. Limonene, for example, a monoterpene found in many citrus oils, can arise from geranyl pyrophosphate by a *cis*-*trans* double-bond isomerization followed by internal nucleophilic displacement of the pyrophosphate group and subsequent loss of a proton.



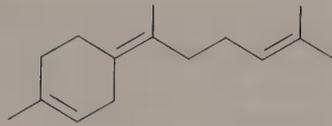
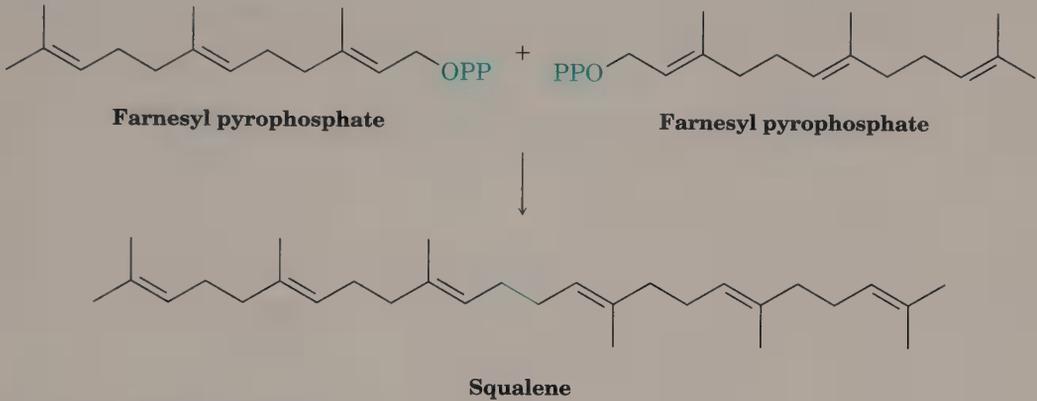
When geranyl pyrophosphate reacts with isopentenyl pyrophosphate, the 15-carbon *farnesyl pyrophosphate* results. Just as geranyl pyrophosphate is the precursor of all monoterpenes, farnesyl pyrophosphate is the precursor of all sesquiterpenes. Farnesol, the corresponding alcohol, is found in citronella oil and lemon oil.



Addition to farnesyl pyrophosphate of further isoprene units gives the 20-carbon and 25-carbon units that serve as precursors of diterpenes and sesterterpenes, respectively. Triterpenes, however, arise biosynthetically by tail-to-tail coupling of two farnesyl pyrophosphates to give *squalene*, a 30-carbon hexaene (Figure 28.6). Squalene, found in high concentration in shark oil, is the precursor from which all other triterpenes and steroids arise.

**PROBLEM** .....

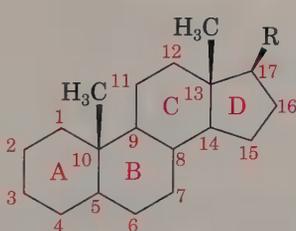
- 28.9** Propose a plausible pathway to account for the biosynthetic formation of  $\gamma$ -bisabolene from farnesyl pyrophosphate.

 $\gamma$ -Bisabolene

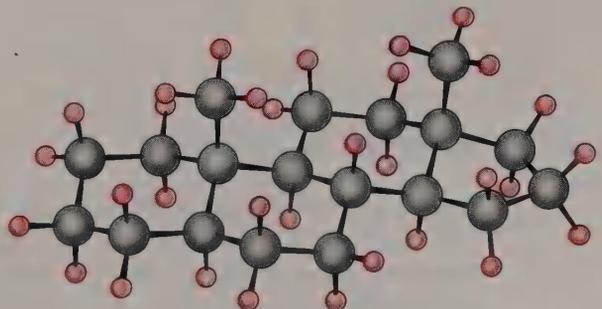
**Figure 28.6** Biosynthesis of squalene, a  $C_{30}$  precursor of triterpenes and steroids, by tail-to-tail coupling of two farnesyl pyrophosphate units.

## 28.8 Steroids

In addition to fats, phospholipids, and terpenes, the lipid extracts of plants and animals also contain steroids. A **steroid** is an organic molecule whose structure is based on the tetracyclic ring system shown below. The four rings are designated A, B, C, and D, beginning at the lower left, and the carbon atoms are numbered beginning in the A ring.



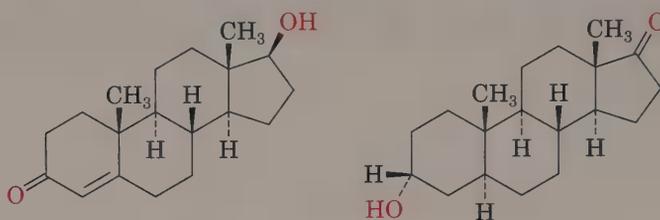
A steroid  
(R = various side chains)



In humans, most steroids function as **hormones**, chemical messengers that are secreted by glands and are carried through the bloodstream to target tissues. There are two main classes of steroid hormones: the *sex hormones*, which control maturation and reproduction, and the *adrenocortical hormones*, which regulate a variety of metabolic processes.

## SEX HORMONES

*Testosterone* and *androsterone* are the two most important male sex hormones, or **androgens**. Androgens are responsible for the development of male secondary sex characteristics during puberty and for promoting tissue and muscle growth. Both are synthesized in the testes from cholesterol.

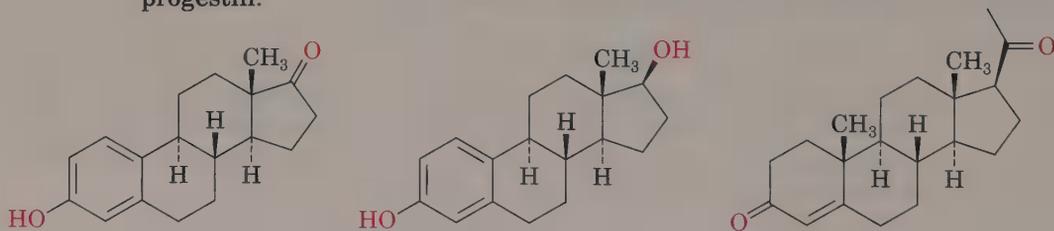


Testosterone

Androsterone

(androgens)

*Estrone* and *estradiol* are the two most important female sex hormones, or **estrogens**. Synthesized in the ovaries from testosterone, estrogenic hormones are responsible for the development of female secondary sex characteristics and for regulation of the menstrual cycle. Note that both have a benzene-like aromatic A ring. In addition, another kind of sex hormone called a *progestin* is essential for preparing the uterus for implantation of a fertilized ovum during pregnancy. *Progesterone* is the most important progestin.



Estrone

Estradiol

Progesterone

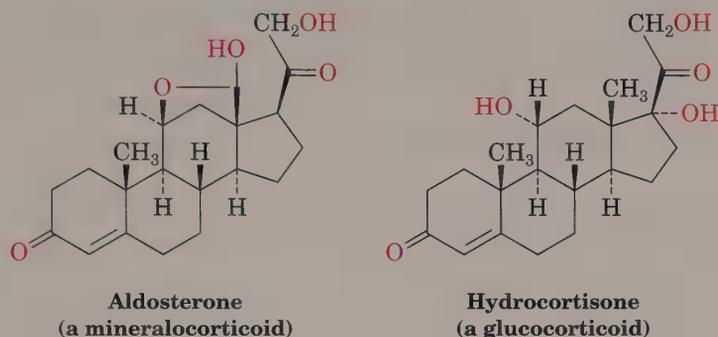
(estrogens)

(a progestin)

## ADRENOCORTICAL HORMONES

Adrenocortical steroids are secreted by the adrenal glands, small organs located near the upper end of each kidney. There are two types of adrenocortical steroids, called *mineralocorticoids* and *glucocorticoids*. Mineralocorticoids such as *aldosterone* control tissue swelling by regulating cellular

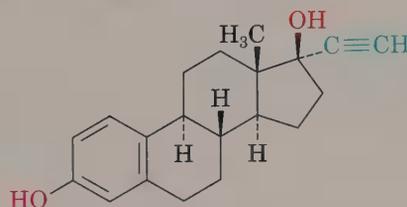
salt balance between  $\text{Na}^+$  and  $\text{K}^+$ . Glucocorticoids such as *hydrocortisone* are involved in the regulation of glucose metabolism and in the control of inflammation. Glucocorticoid ointments are widely used to bring down the swelling from exposure to poison oak or poison ivy.



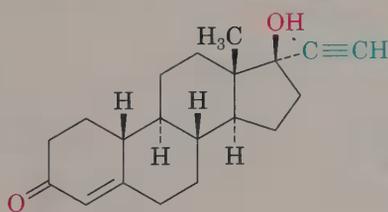
## SYNTHETIC STEROIDS

In addition to the many hundreds of steroids isolated from plants and animals, thousands more have been synthesized in pharmaceutical laboratories in the search for new drugs. The idea is to start with a natural hormone, carry out a chemical modification of the structure, and then see what biological properties the modified steroid has.

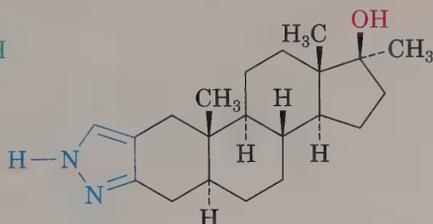
Among the best-known synthetic steroids are oral contraceptive and anabolic agents. Most birth-control pills are a mixture of two compounds, a synthetic estrogen such as *ethynylestradiol* and a synthetic progestin such as *norethindrone*. Anabolic steroids such as *stanozolol*, detected in several athletes during the 1988 Olympics, are synthetic androgens that mimic the tissue-building effects of natural testosterone.



**Ethynylestradiol**  
(a synthetic estrogen)



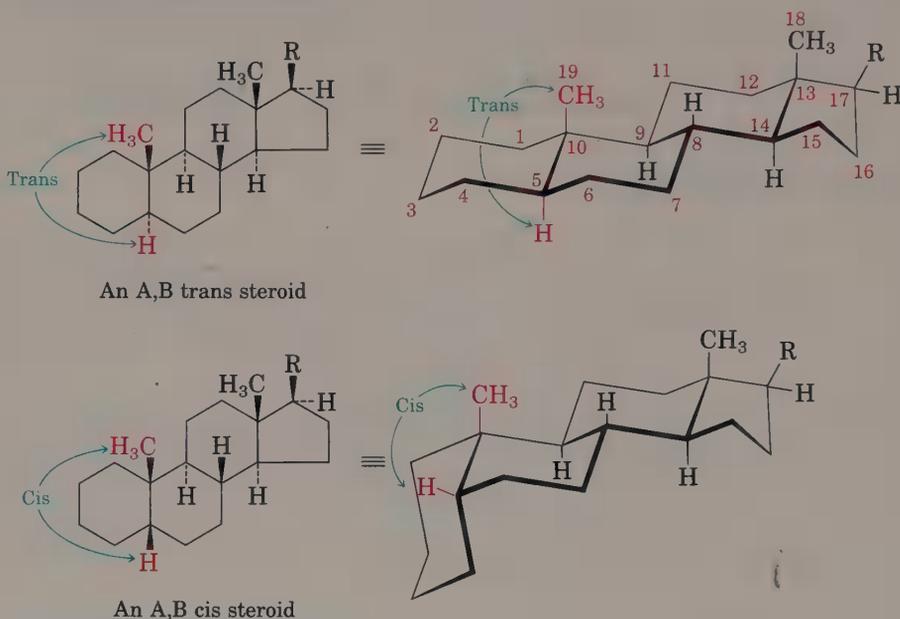
**Norethindrone**  
(a synthetic progestin)



**Stanozolol**  
(an anabolic agent)

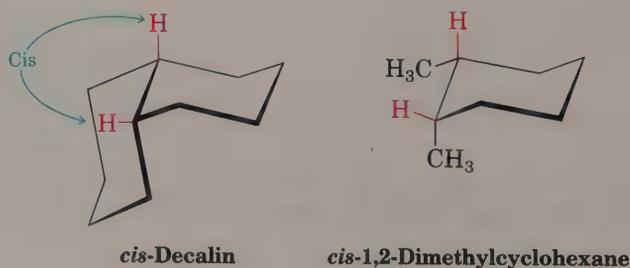
## 28.9 Stereochemistry of Steroids

The steroid skeleton has four rings fused together with a specific stereochemistry. All three of the six-membered rings (rings A, B, and C) can adopt strain-free chair conformations, as indicated in Figure 28.7. Unlike simple cyclohexane rings, however, which can undergo chair-chair interconversions (Section 4.11), steroids are constrained to a rigid conformation and can't undergo ring-flips.



**Figure 28.7** Steroid conformations. The three six-membered rings have chair conformations but are unable to undergo ring-flips.

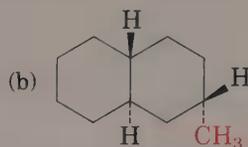
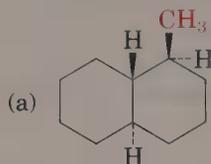
Two cyclohexane rings can be joined in either a *cis* or a *trans* manner. In *cis*-decalin, both groups at the ring-junction positions (the *angular* groups) are on the same side of the two rings. In *trans*-decalin, the groups at the ring junctions are on opposite sides. These spatial relationships are best grasped by building molecular models of *cis*- and *trans*-decalin.





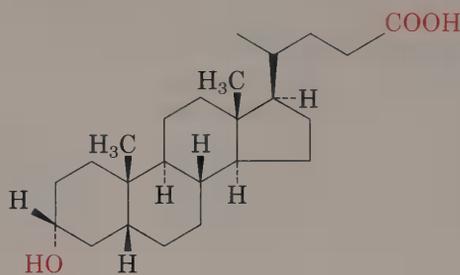
## PROBLEM .....

- 28.10 Draw the following molecules in chair conformations, and indicate whether the ring substituents are axial or equatorial.



## PROBLEM .....

- 28.11 Lithocholic acid is an A-B cis steroid found in human bile. Draw lithocholic acid showing chair conformations as in Figure 28.7, and tell whether the hydroxyl group at C3 is axial or equatorial.



Lithocholic acid

## 28.10 Steroid Biosynthesis

Steroids are heavily modified triterpenes that are biosynthesized in living organisms from the acyclic hydrocarbon squalene (Section 28.7). The exact pathway by which this remarkable transformation is accomplished is lengthy and complex, but most of the key steps have now been worked out, with notable contributions made by Konrad Bloch<sup>4</sup> and John Cornforth,<sup>5</sup> who received Nobel Prizes for their efforts.

Steroid biosynthesis occurs by enzyme-catalyzed epoxidation of squalene to yield squalene oxide, followed by acid-catalyzed cyclization and an extraordinary cascade of multiple carbocation rearrangements to yield lanosterol (Figure 28.9). Lanosterol is then degraded by other enzymes to produce cholesterol, which is itself converted by other enzymes to produce a host of different steroids.

The exact series of carbocation rearrangements involved in the biosynthetic conversion of squalene to lanosterol involves the following steps:

<sup>4</sup>Konrad E. Bloch (1912– ); b. Neisse, Germany; Ph.D. Columbia; professor, University of Chicago, Harvard University; Nobel Prize in medicine (1964).

<sup>5</sup>John Warcup Cornforth (1917– ); b. Australia; Ph.D. Oxford (Robinson); National Institute of Medical Research (Great Britain); Nobel Prize (1975).

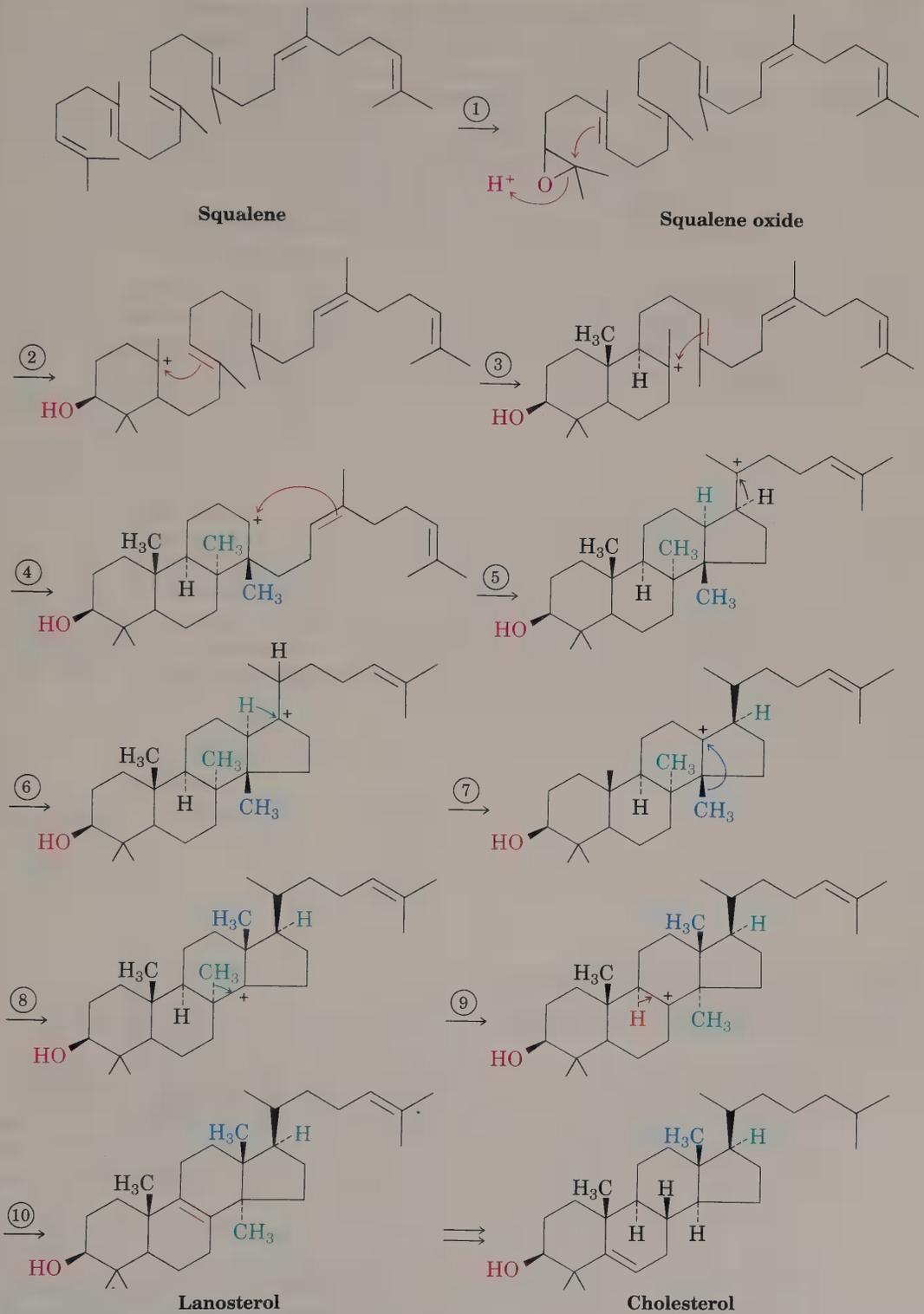


Figure 28.9 Biosynthesis of cholesterol from squalene.

- Step 1* The enzyme *squalene oxidase* selectively epoxidizes a terminal double bond of squalene to yield squalene oxide.
- Step 2* Squalene oxide is protonated, and the epoxide ring is opened by nucleophilic attack of the double bond six carbons away, to yield a cyclic carbocation intermediate.
- Step 3* The tertiary carbocation intermediate produced in step 2 is attacked by another double bond six carbons away to yield a second carbocation intermediate.
- Step 4* A third cyclization occurs by attack of an appropriately positioned double bond on a carbocation.
- Step 5* A fourth and last cyclization takes place, this one giving a five-membered ring.
- Step 6* Carbocation rearrangement occurs by a hydride shift.
- Step 7* A second hydride shift gives still another carbocation.
- Step 8* Carbocation rearrangement occurs by shift of a methyl group.
- Step 9* A second methyl-group shift gives a final carbocation intermediate.
- Step 10* Loss of a proton (E1 reaction) from the carbon next to the cationic center gives lanosterol.

Although written in a stepwise way in Figure 28.9 for convenience, it's thought that the entire cyclization sequence (steps 2–5) takes place at one time without true intermediates being formed. Similarly, the carbocation rearrangements and proton loss (steps 6–10) take place at essentially the same time without intermediates.

PROBLEM.....

- 28.12 Look at the structures of lanosterol and cholesterol, and catalog the changes that have occurred in the transformation.
- .....

## 28.11 Summary and Key Words

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**Lipids** are the naturally occurring materials isolated from plant and animal cells by extraction with organic solvents. They usually have large hydrocarbon portions in their structure. **Animal fats** and **vegetable oils** are the most widely occurring lipids. Both are triesters of glycerol with long-chain **fatty acids**, but animal fats are usually saturated, whereas vegetable oils usually have unsaturated fatty acid residues.

**Phosphoglycerides** such as **lecithin** and **cephalin** are closely related to fats. The glycerol backbone in these molecules is esterified to two fatty acids (one saturated and one unsaturated) and to one phosphate ester. **Sphingolipids**, another major class of **phospholipids**, have an amino alcohol such as sphingosine for their backbone. These compounds are important constituents of cell membranes.

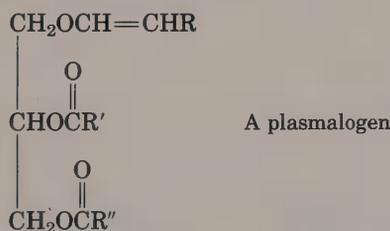
Fatty acids are biosynthesized in nature by condensation of enzyme-bound two-carbon acetate units. The overall scheme is lengthy, but the individual steps are the fundamental organic reactions expected of the functional groups involved.

**Prostaglandins** and **terpenes** are still other classes of lipids. Prostaglandins, which are found in all body tissues, have a wide range of physiological actions. Terpenes are often isolated from the essential oils of plants. They have an immense diversity of structure and are produced biosynthetically by head-to-tail coupling of the five-carbon "isoprene equivalents," isopentenyl pyrophosphate and dimethylallyl pyrophosphate.

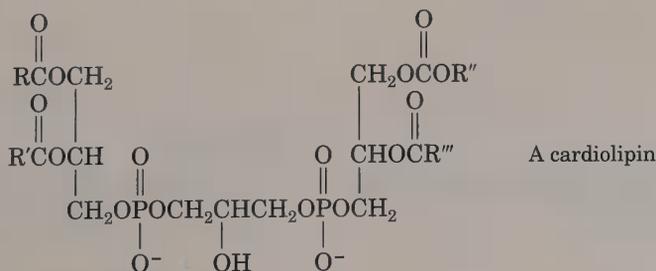
**Steroids** are plant and animal lipids with a characteristic tetracyclic carbon skeleton. Like the prostaglandins, steroids occur widely in body tissue and have a large variety of physiological activities. Steroids are closely related to terpenes and arise biosynthetically from the triterpene precursor, lanosterol. Lanosterol, in turn, arises from cyclization of the acyclic hydrocarbon, squalene.

## ADDITIONAL PROBLEMS .....

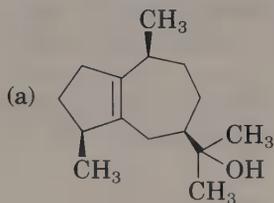
- 28.13 Fats can be either optically active or optically inactive, depending on their structure. Draw the structure of an optically active fat that yields two equivalents of stearic acid and one equivalent of oleic acid on hydrolysis. Draw the structure of an optically inactive fat that yields the same products.
- 28.14 Spermaceti, a fragrant substance from sperm whales, was much used in cosmetics until it was banned in 1976 to protect the whales from extinction. Chemically, spermaceti is cetyl palmitate, the ester of cetyl alcohol ( $n\text{-C}_{16}\text{H}_{33}\text{OH}$ ) with palmitic acid. Draw its structure.
- 28.15 The *plasmalogens* are a group of lipids found in nerve and muscle cells. Look at the following plasmalogen structure, and tell how it differs from fats, lecithins, and cephalins.



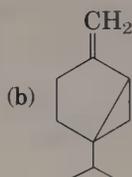
- 28.16 What products would you obtain from hydrolysis of a plasmalogen (Problem 28.15) with aqueous NaOH? With  $\text{H}_3\text{O}^+$ ?
- 28.17 *Cardiolipins* are a group of lipids found in heart muscles. What products would be formed if all ester bonds were saponified by treatment with aqueous NaOH?



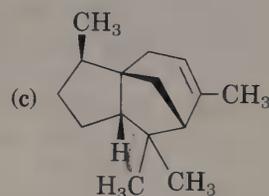
- 28.18 Show the products you would expect to obtain from reaction of glyceryl trioleate with the following:
- (a) Excess  $\text{Br}_2$  in  $\text{CCl}_4$  (b)  $\text{H}_2/\text{Pd}$   
 (c)  $\text{NaOH}/\text{H}_2\text{O}$  (d)  $\text{O}_3$ , then  $\text{Zn}/\text{CH}_3\text{COOH}$   
 (e)  $\text{LiAlH}_4$ , then  $\text{H}_3\text{O}^+$  (f)  $\text{CH}_3\text{MgBr}$ , then  $\text{H}_3\text{O}^+$   
 (g)  $\text{NaOH}/\text{H}_2\text{O}$ , then  $\text{CH}_2\text{N}_2$  (diazomethane)
- 28.19 Eleostearic acid,  $\text{C}_{18}\text{H}_{30}\text{O}_2$ , is a rare fatty acid found in the tung oil used for furniture finishing. On ozonolysis followed by treatment with zinc, eleostearic acid furnishes one part pentanal, two parts glyoxal ( $\text{OHC}-\text{CHO}$ ), and one part 9-oxononanoic acid [ $\text{OHC}(\text{CH}_2)_7\text{COOH}$ ]. What is the structure of eleostearic acid?
- 28.20 Draw representative structures for:
- (a) A fat (b) A prostaglandin (c) A steroid
- 28.21 How would you convert oleic acid into these substances?
- (a) Methyl oleate (b) Methyl stearate  
 (c) Nonanal (d) Nonanedioic acid  
 (e) 9-Octadecynoic acid (stearolic acid) (f) 2-Bromostearic acid  
 (g) 18-Pentatriacontanone,  $\text{CH}_3(\text{CH}_2)_{16}\text{CO}(\text{CH}_2)_{16}\text{CH}_3$
- 28.22 How would you synthesize stearolic acid (Problem 28.21e) from 1-decyne and 1-chloro-7-iodoheptane?
- 28.23 Vaccenic acid,  $\text{C}_{18}\text{H}_{34}\text{O}_2$ , is a rare fatty acid that gives heptanal and 11-oxoundecanoic acid [ $\text{OHC}(\text{CH}_2)_9\text{COOH}$ ] on ozonolysis followed by zinc treatment. When allowed to react with  $\text{CH}_2\text{I}_2/\text{Zn}-\text{Cu}$ , vaccenic acid is converted into lactobacillic acid. What are the structures of vaccenic and lactobacillic acids?
- 28.24 Show the location of the isoprene units in these terpenes:



Guaiol



Sabinene

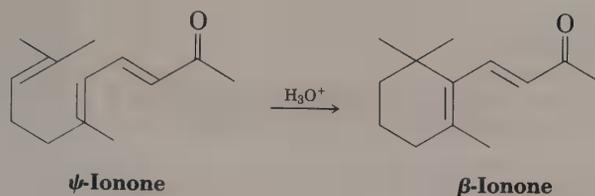


Cedrene

- 28.25 Indicate by asterisks the stereogenic centers present in each of the three terpenes shown in Problem 28.24. How many stereoisomers of each are theoretically possible?
- 28.26 Assume that the three terpenes in Problem 28.24 are derived biosynthetically from isopentenyl pyrophosphate and dimethylallyl pyrophosphate, each of which was isotopically labeled at the pyrophosphate-bearing carbon atom (C1). At what positions would the terpenes be isotopically labeled?
- 28.27 Suggest a mechanistic pathway by which  $\alpha$ -pinene might arise biosynthetically from geranyl pyrophosphate.

 $\alpha$ -Pinene

- 28.28 Suggest a mechanism by which  $\psi$ -ionone is transformed into  $\beta$ -ionone on treatment with acid.

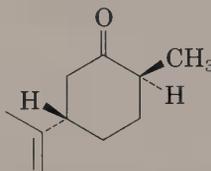


- 28.29 Which isomer would you expect to be more stable, *cis*-decalin or *trans*-decalin? Explain. (You might want to review Section 4.13.)



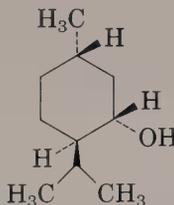
Decalin

- 28.30 Draw the most stable chair conformation of dihydrocarvone.



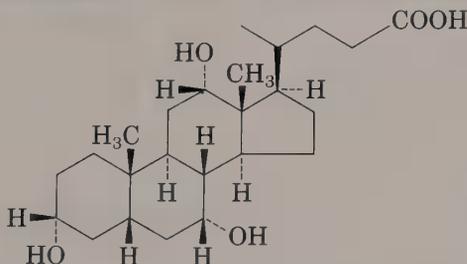
Dihydrocarvone

- 28.31 Draw the most stable chair conformation of menthol, and label each substituent as axial or equatorial.



Menthol (from peppermint oil)

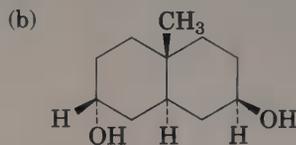
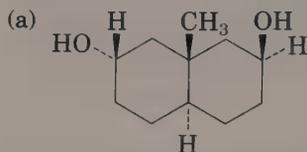
- 28.32 Cholic acid, a major steroidal constituent of human bile, has the structure shown. Draw a conformational structure of cholic acid, and label the three hydroxyl groups as axial or equatorial.



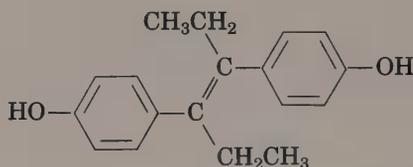
Cholic acid

- 28.33 How many stereogenic centers does cholic acid (Problem 28.32) have? How many stereoisomers are possible?

- 28.34 Show the products you would expect to obtain from reaction of cholic acid with these reagents:
- $C_2H_5OH, HCl$
  - Excess pyridinium chlorochromate in  $CH_2Cl_2$
  - $BH_3$  in THF, then  $H_3O^+$
- 28.35 As a general rule, equatorial alcohols are esterified more readily than axial alcohols. What product would you expect to obtain from reaction of these two compounds with one equivalent of acetic anhydride?

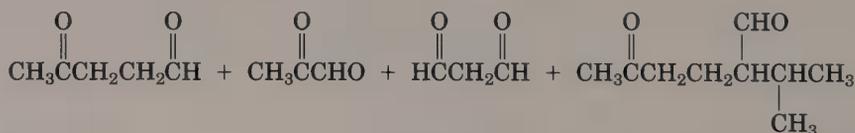


- 28.36 Diethylstilbestrol (DES) exhibits estradiol-like activity even though it is structurally unrelated to steroids. Once used widely as an additive in animal feed, DES has been implicated as a causative agent in several types of cancer. Look up the structure of estradiol (Section 28.8), and show how DES can be drawn so that it is sterically similar to estradiol.

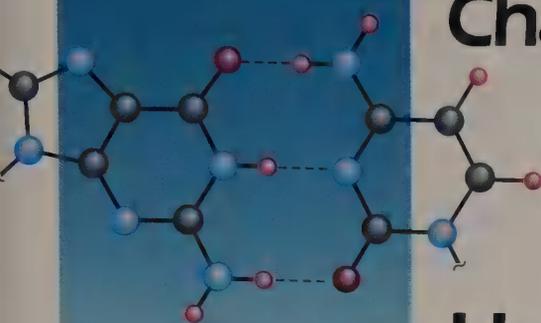


Diethylstilbestrol

- 28.37 Propose a synthesis of diethylstilbestrol (Problem 28.36) from phenol and any other organic compound required.
- 28.38 What products would you expect from reaction of estradiol (Problem 28.36) with these reagents?
- $NaH$ , then  $CH_3I$
  - $CH_3COCl$ , pyridine
  - $Br_2, FeBr_3$
  - Pyridinium chlorochromate in  $CH_2Cl_2$
- 28.39 Cembrene,  $C_{20}H_{32}$ , is a diterpene hydrocarbon isolated from pine resin. Cembrene has an ultraviolet absorption at 245 nm, but dihydrocembrene ( $C_{20}H_{34}$ ), the product of hydrogenation with one equivalent of hydrogen, has no ultraviolet absorption. On exhaustive hydrogenation, four equivalents of hydrogen react, and octahydrocembrene,  $C_{20}H_{40}$ , is produced. On ozonolysis of cembrene, followed by treatment of the ozonide with zinc, four carbonyl-containing products are obtained:



Propose a structure for cembrene that is consistent with the isoprene rule.

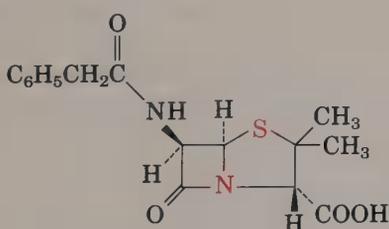


# Chapter 29

## Heterocycles and Nucleic Acids

Cyclic organic compounds are classified either as **carbocycles** or as **heterocycles**. Carbocyclic rings contain only carbon atoms, but heterocyclic rings contain one or more different atoms in addition to carbon. Nitrogen, oxygen, and sulfur are the most common heteroatoms, but many others are also known.

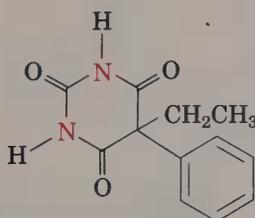
Heterocyclic compounds are common in organic chemistry, and many have important biological properties. For example, the antibiotic penicillin, the antiulcer agent cimetidine, the sedative phenobarbital, and the non-nutritive sweetener saccharin are all heterocycles:



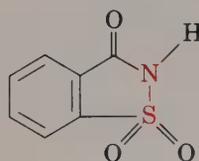
Penicillin G



Cimetidine



Phenobarbital



Saccharin





All three of these heterocycles are liquid at room temperature, as indicated by the data in Table 29.1.

**TABLE 29.1** Physical Properties of Some Five-Membered Heterocycles

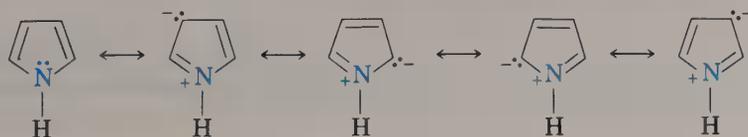
Name	Molecular weight	Melting point ( $^\circ\text{C}$ )	Boiling point ( $^\circ\text{C}$ )
Furan	68	-85	31
Pyrrole	67	-23	130
Thiophene	84	-38	

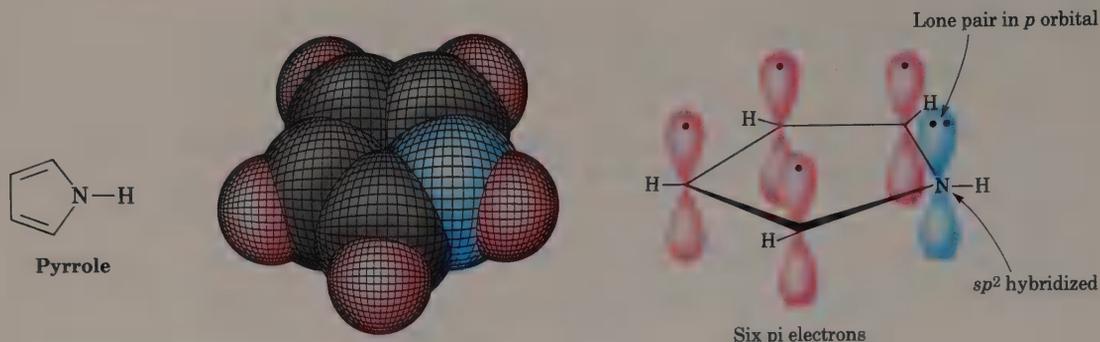
The chemistry of all three heterocyclic ring systems contains some surprises. For example, pyrrole is both an amine and a conjugated diene, yet its chemical properties aren't consistent with either of these structural features. Unlike most other amines, pyrrole is not basic; unlike most other conjugated dienes, pyrrole undergoes electrophilic substitution rather than addition reactions. The same is true of furan and thiophene: Both react with electrophiles to give substitution products.

## 29.2 Structures of Pyrrole, Furan, and Thiophene

Pyrrole, furan, and thiophene give electrophilic substitution products because they're *aromatic*. Each has six pi electrons ( $4n + 2$ , where  $n = 1$ ) in a cyclic conjugated system. To choose pyrrole as an example, each of the four carbon atoms of pyrrole contributes one pi electron, and the  $sp^2$ -hybridized nitrogen atom contributes two (its lone pair). The six pi electrons occupy  $p$  orbitals, with lobes above and below the plane of the ring, as shown in Figure 29.1. Overlap of the five  $p$  orbitals forms aromatic molecular orbitals just as in benzene (Section 15.9).

Note that the pyrrole nitrogen atom uses all five of its valence electrons in bonding. Three electrons are used in forming three sigma bonds (two to carbon and one to hydrogen), and the two lone-pair electrons are involved in aromatic pi bonding, as the following resonance structures indicate:





**Figure 29.1** Pyrrole, a six-pi-electron aromatic heterocycle.

Because the nitrogen lone pair is a part of the aromatic sextet, it is less available for bonding to electrophiles, and pyrrole is therefore less basic ( $pK_a$  of pyrrolinium ion  $\approx 0$ ) and less nucleophilic than aliphatic amines. By the same token, however, the carbon atoms of pyrrole are *more* electron-rich and more nucleophilic than typical double-bond carbon atoms. The pyrrole ring is therefore reactive toward electrophiles in the same way that activated benzene rings are reactive.

**PROBLEM** .....

- 29.1 Draw an orbital picture of furan. Assume that the oxygen atom is  $sp^2$  hybridized, and show the orbitals that the two oxygen lone pairs are occupying.

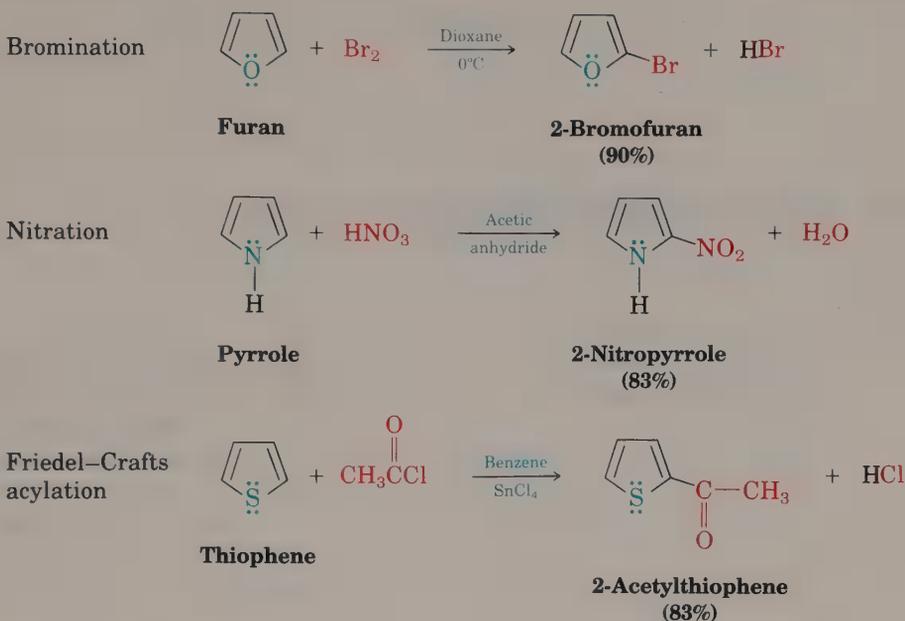
**PROBLEM** .....

- 29.2 Pyrrole has a dipole moment of 1.8 D. Look at the several possible resonance structures of pyrrole, and predict which is the positive end and which is the negative end of the dipole.
- .....

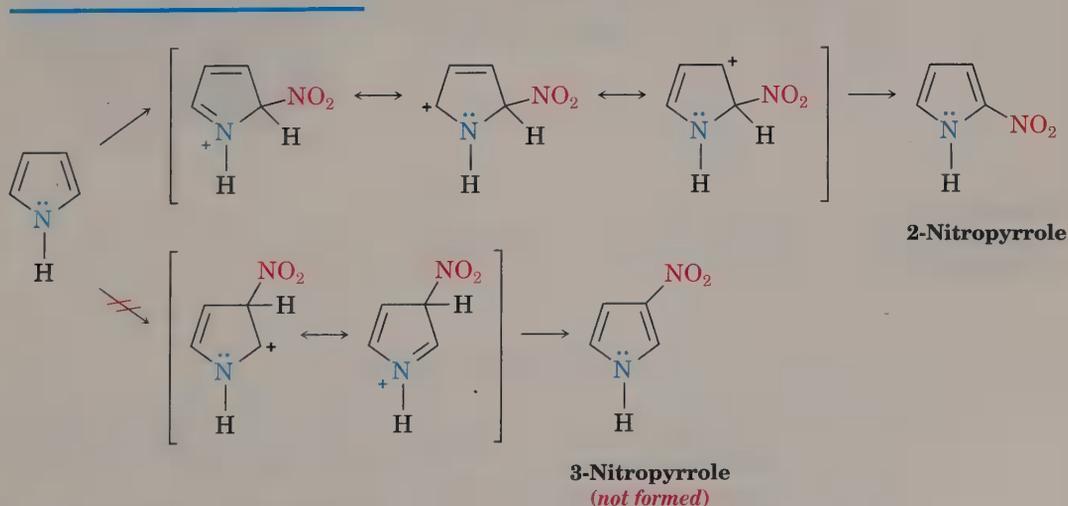
### 29.3 Electrophilic Substitution Reactions of Pyrrole, Furan, and Thiophene

The chemistry of pyrrole, furan, and thiophene is similar to that of activated benzenoid aromatic rings. Like benzene, the five-membered aromatic heterocycles undergo electrophilic substitution rather than addition reactions. In general, these heterocycles are much more reactive toward electrophiles than benzene rings are, and low temperatures are often necessary to prevent destruction of the starting material. Halogenation, nitration, sulfonation, and Friedel–Crafts acylation can all be accomplished if the proper reaction

conditions are chosen. A reactivity order of furan > pyrrole > thiophene is normally found.



Electrophilic substitution normally occurs at C2, the position next to the heteroatom, because C2 is the most electron-rich, most nucleophilic position on the ring. Another way of saying the same thing is to note that electrophilic attack at C2 leads to a more stable intermediate cation, with three resonance forms, than does attack at C3, which results in a cation with only two resonance forms (Figure 29.2).



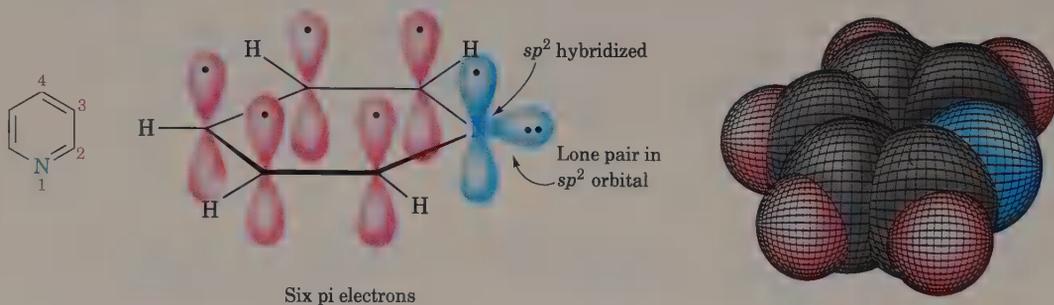
**Figure 29.2** Electrophilic nitration of pyrrole. The intermediate produced by reaction at C2 is more stable than that produced by reaction at C3.

## PROBLEM .....

- 29.3 Propose a mechanism to account for the fact that treatment of pyrrole with deuterio-sulfuric acid,  $D_2SO_4$ , leads to formation of 2-deuteriopyrrole.
- .....

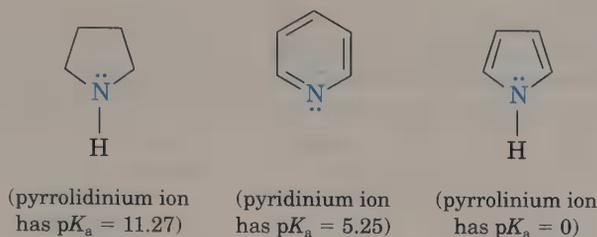
## 29.4 Pyridine, a Six-Membered Heterocycle

Pyridine, obtained commercially by distillation of coal tar, is the nitrogen-containing heterocyclic analog of benzene. Like benzene, pyridine is aromatic. It is a flat molecule with bond angles of  $120^\circ$  and with carbon-carbon bond lengths of  $1.39 \text{ \AA}$ , intermediate between normal single and double bonds. Each of the five carbon atoms contributes one pi electron, and the  $sp^2$ -hybridized nitrogen atom also contributes one pi electron to complete the aromatic sextet. Unlike the situation in pyrrole, however, the lone pair of electrons on the pyridine nitrogen atom is not involved in bonding but occupies an  $sp^2$  orbital in the plane of the ring (Figure 29.3).



**Figure 29.3** Electronic structure of pyridine, a six-pi-electron nitrogen-containing analog of benzene.

Pyridine is a stronger base than pyrrole ( $pK_a = 5.25$  for pyridinium ion versus 0 for pyrrolinium ion) because of its electronic structure. Since the pyridine nitrogen's lone-pair electrons aren't involved in aromatic pi bonding, they're available for donation to a Lewis acid. Pyridine is a weaker base than aliphatic amines, however ( $pK_a = 5.25$  versus 11), because of its hybridization. The lone-pair electrons in the  $sp^3$ -hybridized nitrogen orbital ( $\frac{1}{4}$  s character) of an aliphatic amine are held less closely to the positively charged nucleus than are lone-pair electrons in an  $sp^2$  hybrid orbital ( $\frac{1}{3}$  s character), with the result that  $sp^3$ -hybridized nitrogen is more basic.



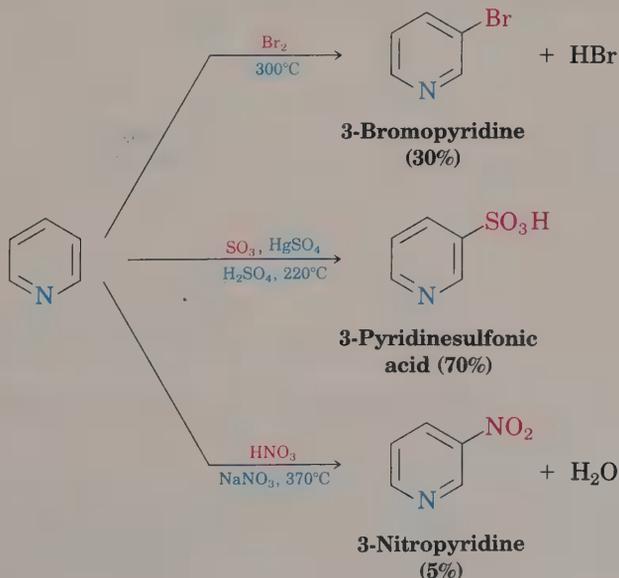
## PROBLEM

- 29.4 Imidazolium ion has  $pK_a = 6.95$ . Draw an orbital structure of imidazole, and indicate which nitrogen is more basic.

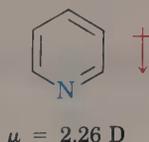


## 29.5 Electrophilic Substitution of Pyridine

The pyridine ring undergoes electrophilic aromatic substitution reactions with great difficulty. Halogenation and sulfonation can be carried out under drastic conditions, but nitration occurs in very low yield, and Friedel–Crafts reactions aren't successful. Reactions usually give the 3-substituted product.



The low reactivity of pyridine toward electrophilic aromatic substitution is due to a combination of factors. Most important is that the electron density of the ring is decreased by the electron-withdrawing inductive effect of the electronegative nitrogen atom. Thus, pyridine has a dipole moment of 2.26 D, with the ring acting as the positive end of the dipole. Electrophilic attack on the positively polarized ring is therefore difficult.



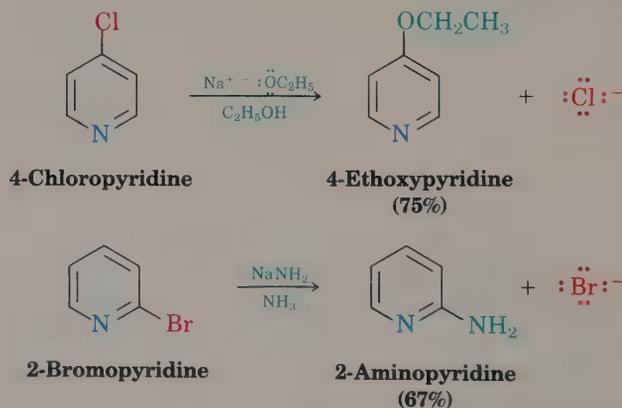
A second factor decreasing the reactivity of the pyridine ring toward electrophilic attack is that acid–base complexation between the basic ring nitrogen atom and the attacking electrophile places a positive charge on the ring, further deactivating it.

**PROBLEM** .....

- 29.5** Electrophilic aromatic substitution reactions of pyridine normally occur at C3. Draw the carbocation intermediates resulting from electrophilic attack at all possible positions, and explain the observed result.
- .....

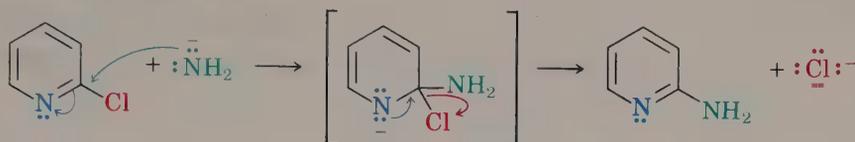
## 29.6 Nucleophilic Substitution of Pyridine

In contrast to their lack of reactivity toward electrophilic substitution, certain substituted pyridine rings undergo *nucleophilic* aromatic substitution with relative ease. Both 2- and 4-halo-substituted (but not 3-substituted) pyridines react particularly well.

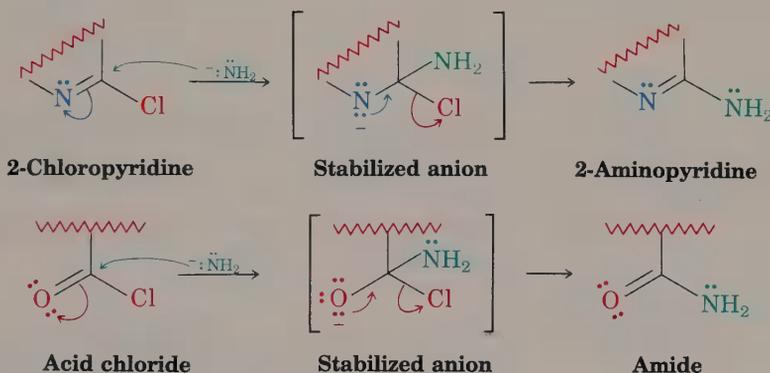


These reactions are typical nucleophilic aromatic substitutions, analogous to those we saw earlier for halo-substituted benzenes (Section 16.9). Although a benzene ring needs to be further activated by the presence of

electron-withdrawing substituents for nucleophilic substitution to occur, pyridines are already sufficiently activated. Reaction occurs by addition of the nucleophile to the C=N bond, followed by loss of halide ion from the anion intermediate.



This nucleophilic aromatic substitution is in some ways analogous to the nucleophilic acyl substitution of acid chlorides (Section 21.5). In both cases, the initial addition step is favored by the ability of the electronegative atom (nitrogen or oxygen) to stabilize the anion intermediate. The intermediate then expels chloride ion to yield the substitution product.



PROBLEM.....

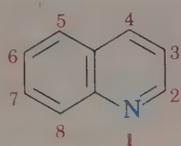
- 29.6 Draw the anion intermediates expected from nucleophilic attack at C4 of a 4-halopyridine and at C3 of a 3-halopyridine. How can you account for the fact that substitution of the 4-halopyridine occurs readily, but the 3-halopyridine does not react?

PROBLEM.....

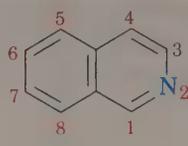
- 29.7 If 3-bromopyridine is heated with  $\text{NaNH}_2$  under forcing conditions, a mixture of 3- and 4-aminopyridine is obtained. Explain. [Hint: See Section 16.10.]
- .....

## 29.7 Fused-Ring Heterocycles

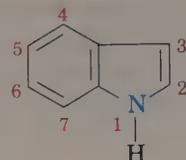
Quinoline, isoquinoline, and indole are **fused-ring heterocycles** that contain both a benzene ring and a heterocyclic aromatic ring. All three ring systems occur commonly in natural products, and many members of the class have pronounced biological activity. Thus, the quinoline alkaloid quinine is widely used as an antimalarial drug, and the indole alkaloid *N,N*-dimethyltryptamine is a powerful hallucinogen.



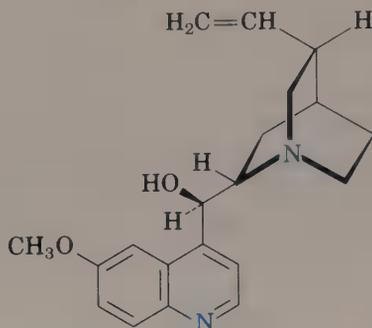
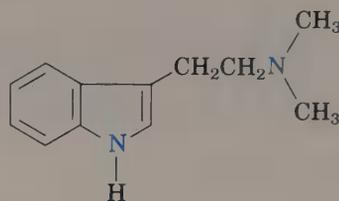
Quinoline



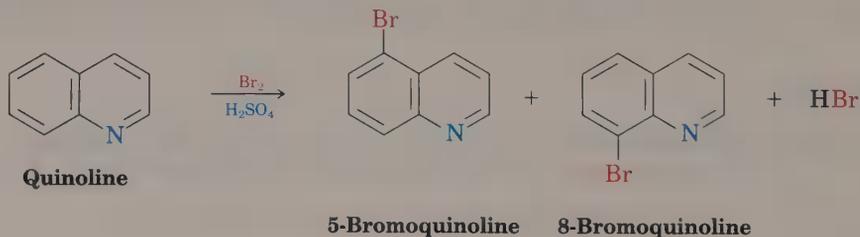
Isoquinoline



Indole

Quinine, an antimalarial drug  
(a quinoline alkaloid)*N,N*-Dimethyltryptamine, a hallucinogen  
(an indole alkaloid)

The chemistry of these three classes of fused-ring heterocycles is just what we might expect from our knowledge of the simpler heterocycles, pyridine and pyrrole. All three undergo electrophilic aromatic substitution reactions. Quinoline and isoquinoline both undergo electrophilic substitution more easily than pyridine but less easily than benzene, consistent with the previous observation that pyridine rings are deactivated compared to benzene. Note that reaction occurs on the benzene ring rather than on the pyridine ring and that a mixture of C5 and C8 substitution products are obtained.

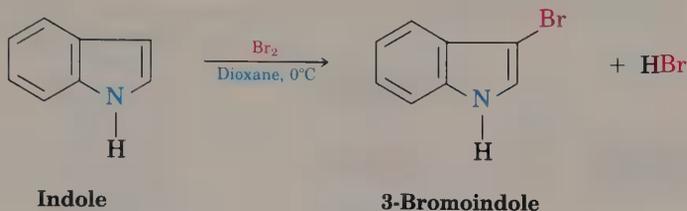


A 51:49 ratio



A 90:10 ratio

Indole undergoes electrophilic substitution more easily than benzene but less easily than pyrrole. Again, this is consistent with the observation that pyrrole rings are more strongly activated than benzene rings. Substitution occurs at C3 of the electron-rich pyrrole ring, rather than on the benzene ring, but reaction conditions must be chosen carefully to avoid destructive side reactions.



PROBLEM .....

29.8 Which nitrogen atom in *N,N*-dimethyltryptamine is more basic? Explain.

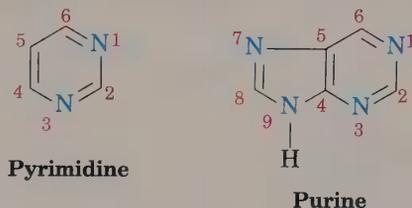
PROBLEM .....

29.9 Indole reacts with electrophiles at C3 rather than at C2. Draw resonance forms of the intermediate cations resulting from attack at C2 and C3, and explain the observed results.

.....

## 29.8 Pyrimidine and Purine

The most important heterocyclic ring systems from a biological viewpoint are *pyrimidine* and *purine*. **Pyrimidine** contains two nitrogens in a six-membered aromatic ring, while **purine** has four nitrogens in a fused-ring structure:



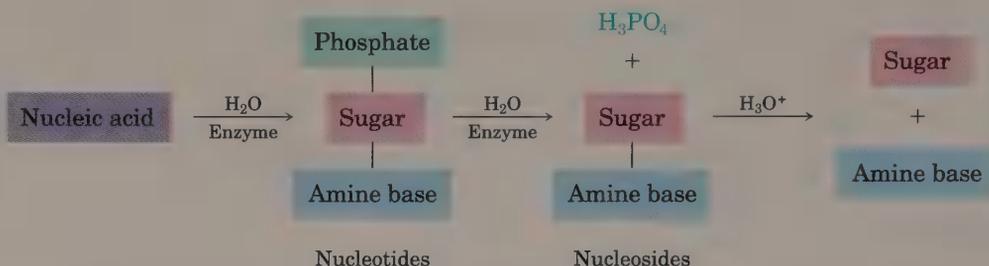
Both heterocycles are essential components of the last major class of biomolecules we'll consider: the nucleic acids.

## 29.9 Nucleic Acids and Nucleotides

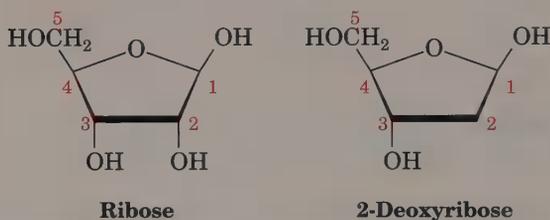
The nucleic acids, **deoxyribonucleic acid (DNA)** and **ribonucleic acid (RNA)**, are the chemical carriers of a cell's genetic information. Coded in a cell's DNA is all the information that determines the nature of the cell,

controls cell growth and division, and directs biosynthesis of the enzymes and other proteins required for all cellular functions.

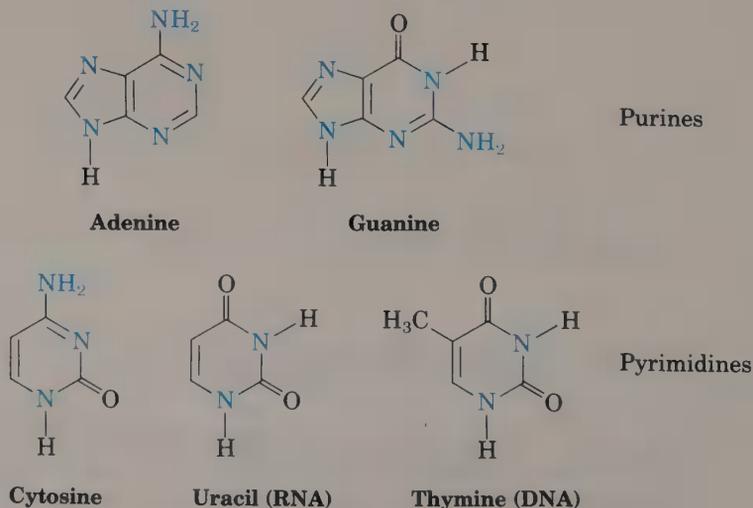
Just as proteins are polymers made of amino acid units, nucleic acids are polymers made of individual building blocks called **nucleotides** that are linked together to form a long chain. Each nucleotide can be cleaved by enzyme-catalyzed hydrolysis to give a **nucleoside** plus phosphoric acid,  $\text{H}_3\text{PO}_4$ , and each nucleoside can be hydrolyzed to yield a simple pentose sugar plus a heterocyclic purine or pyrimidine base.



The sugar component in RNA is ribose, and the sugar in DNA is 2'-deoxyribose. The prefix *2'-deoxy* indicates that oxygen is missing from the 2' position of ribose; numbers with a prime superscript refer to positions on the sugar component of a nucleotide, and numbers without a prime refer to positions on the heterocyclic amine base.

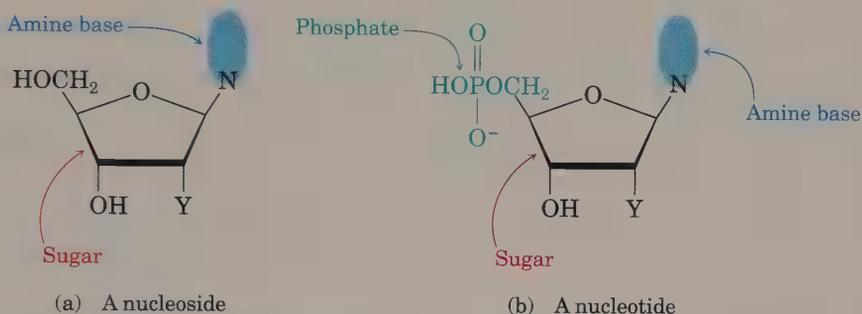


There are four different heterocyclic amine bases in deoxyribonucleotides. Two are substituted purines (**adenine** and **guanine**), and two are substituted pyrimidines (**cytosine** and **thymine**). Adenine, guanine, and



cytosine also occur in RNA, but thymine is replaced in RNA by a different pyrimidine base called **uracil**.

In both DNA and RNA, the heterocyclic amine base is bonded to C1' of the sugar, and the phosphoric acid is bound by a phosphate ester linkage to the C5' sugar position. Thus, nucleosides and nucleotides have the general structure shown in Figure 29.4.



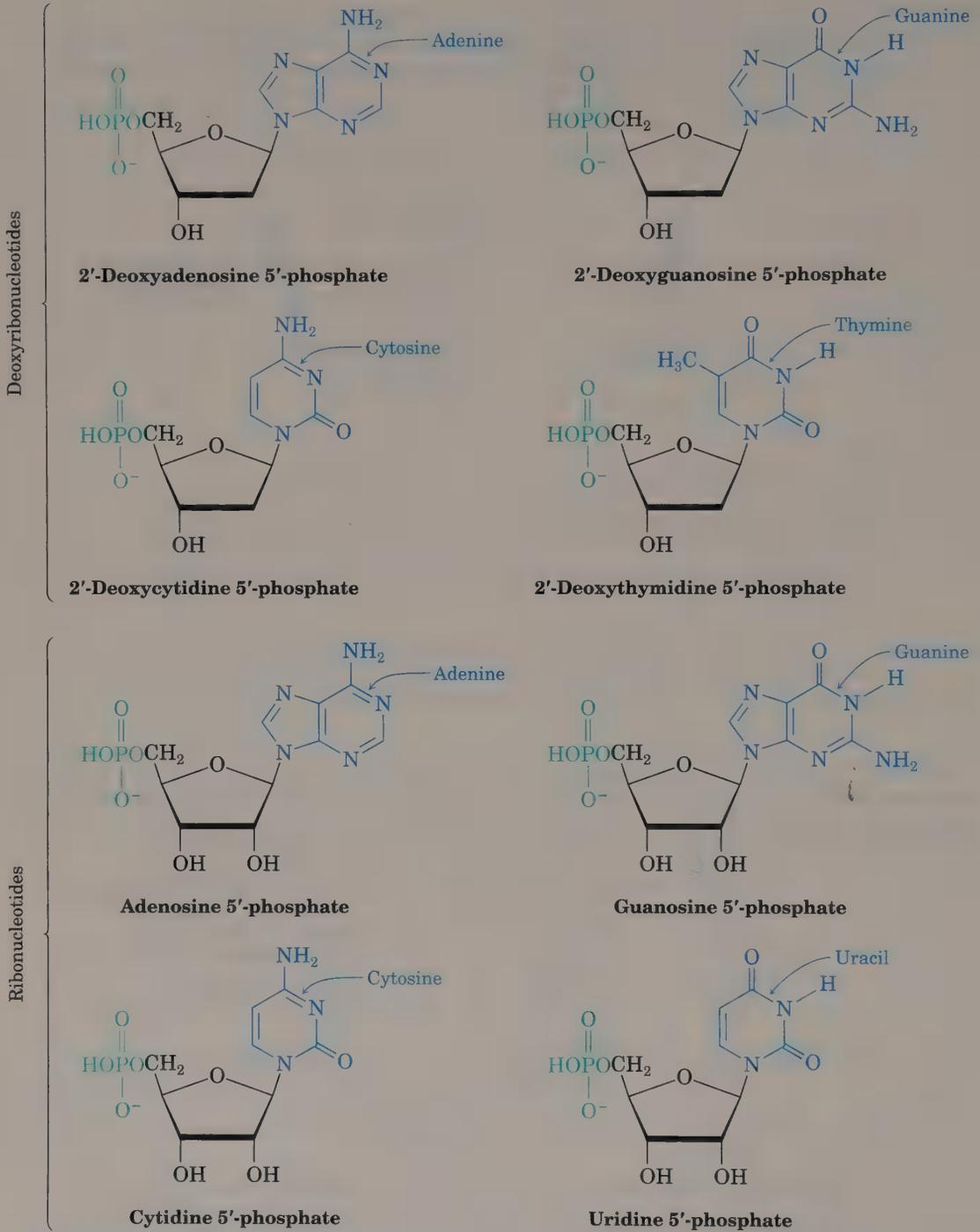
**Figure 29.4** (a) A nucleoside, and (b) a nucleotide. When Y = H, the sugar is deoxyribose. When Y = OH, the sugar is ribose.

The complete names of the bases, nucleotides, and corresponding nucleosides are given in Table 29.2, and the complete structures of all four deoxyribonucleotides and all four ribonucleotides are shown in Figure 29.5.

**TABLE 29.2** Names of Bases, Nucleosides, and Nucleotides

Heterocyclic base	Source	Nucleoside	Nucleotide
Adenine	RNA	Adenosine	Adenosine 5'-phosphate
	DNA	2'-Deoxyadenosine	2'-Deoxyadenosine 5'-phosphate
Guanine	RNA	Guanosine	Guanosine 5'-phosphate
	DNA	2'-Deoxyguanosine	2'-Deoxyguanosine 5'-phosphate
Cytosine	RNA	Cytidine	Cytidine 5'-phosphate
	DNA	2'-Deoxycytidine	2'-Deoxycytidine 5'-phosphate
Uracil	RNA	Uridine	Uridine 5'-phosphate
Thymine	DNA	2'-Deoxythymidine	2'-Deoxythymidine 5'-phosphate

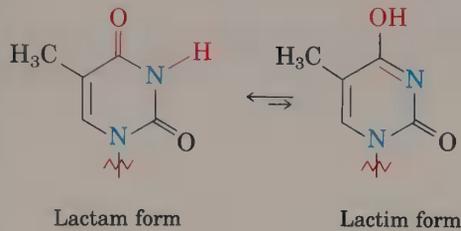
Though chemically similar, DNA and RNA are different in size and have different roles within the cell. Molecules of DNA are enormous, with molecular weights of up to 150 *billion* and lengths up to 12 cm, and are found mostly in the nucleus of the cell. Molecules of RNA, by contrast, are much smaller (as low as 35,000 mol wt) and are found mostly outside the cell nucleus. We'll consider the two kinds of nucleic acids separately, beginning with DNA.



**Figure 29.5** Structures of the four deoxyribonucleotides and the four ribonucleotides.

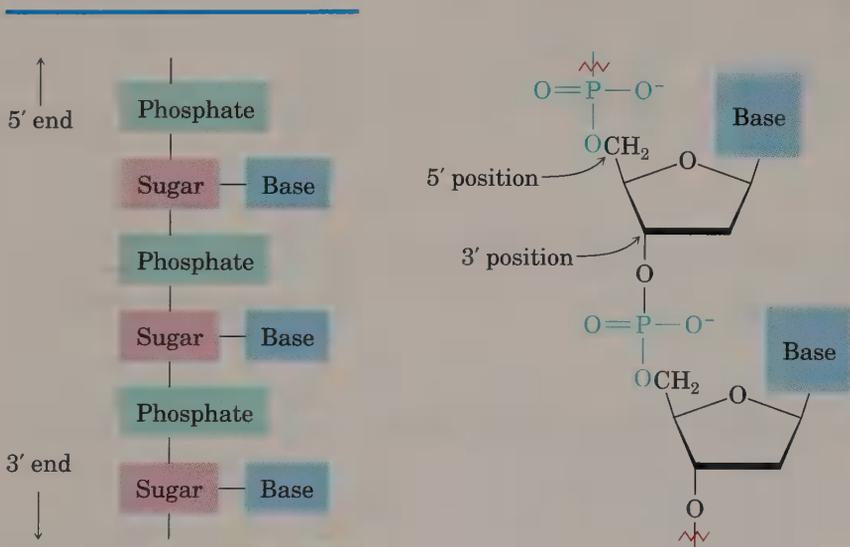
## PROBLEM

- 29.10 2'-Deoxythymidine exists largely in the lactam form rather than in the tautomeric lactim form. Explain.



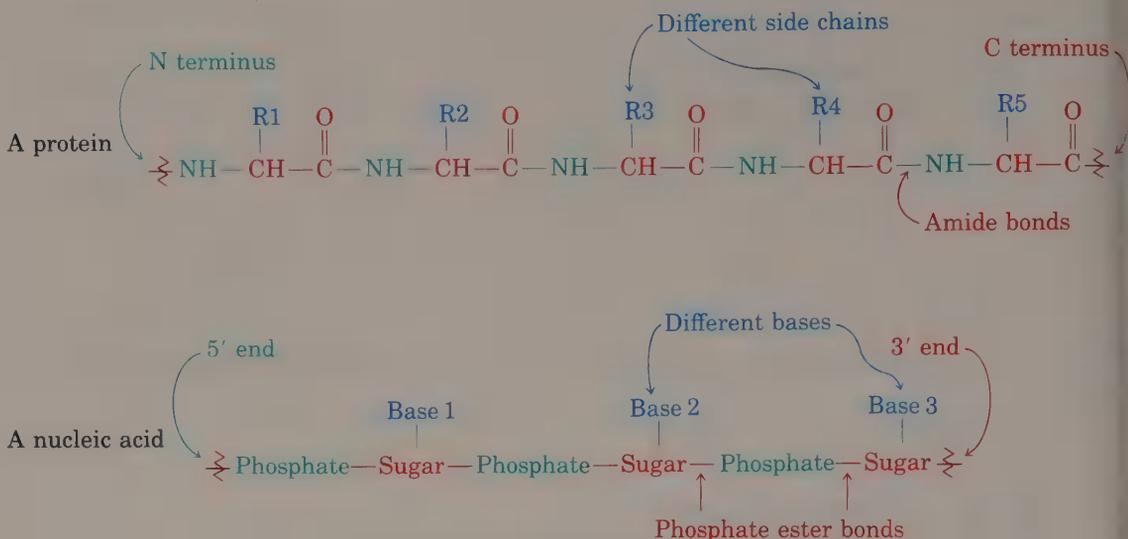
## 29.10 Structure of DNA

Nucleotides join together in DNA by forming a phosphate ester bond between the 5'-phosphate component of one nucleotide and the 3' hydroxyl on the sugar component of another nucleotide (Figure 29.6). One end of the nucleic acid polymer has a free hydroxyl at C3' (called the **3' end**), and the other end has a phosphoric acid residue at C5' (the **5' end**).



**Figure 29.6** Generalized structure of DNA.

Just as the structure of a protein depends on the sequence in which individual amino acid residues are connected, the structure of a nucleic acid depends on the sequence of individual nucleotides. To carry the analogy further, just as a protein has a polyamide backbone with different side chains attached to it, a nucleic acid has an alternating sugar-phosphate backbone with different amine bases attached.



The sequence of nucleotides in a chain is described by starting at the 5' end and identifying the bases in order of occurrence. Rather than write the full name of each nucleotide, though, it's more convenient to use simple abbreviations—A for adenosine, T for thymine, G for guanosine, and C for cytidine. Thus, a typical DNA sequence might be written as -T-A-G-G-C-T-.

**PROBLEM** .....

**29.11** Draw the full structure of the DNA dinucleotide A-G.

**PROBLEM** .....

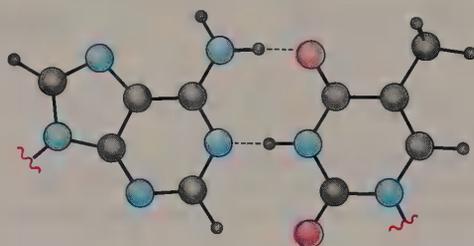
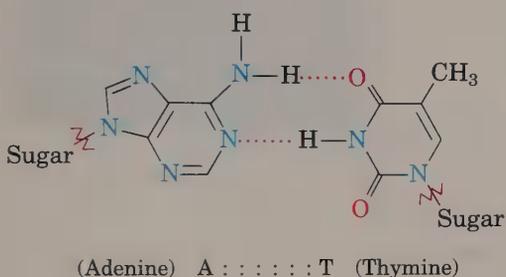
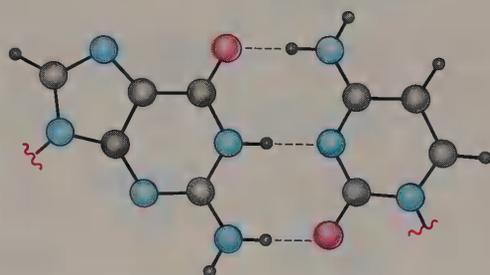
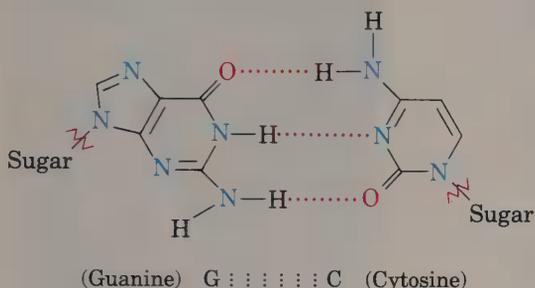
**29.12** Draw the full structure of the RNA dinucleotide U-A.

.....

## 29.11 Base Pairing in DNA: The Watson-Crick Model

Samples of DNA isolated from different tissues of the same species have the same proportions of heterocyclic bases, but samples from different species can have greatly differing proportions of bases. For example, human DNA contains about 30% each of adenine and thymine and about 20% each of guanine and cytosine. The bacterium *Clostridium perfringens*, however, contains about 37% each of adenine and thymine and only 13% each of guanine and cytosine. Note that in both of these examples the bases occur in pairs. Adenine and thymine are usually present in equal amounts, as are cytosine and guanine. Why should this be?

In 1953, James Watson<sup>1</sup> and Francis Crick<sup>2</sup> made their now classic proposal for the secondary structure of DNA. According to the Watson–Crick model, DNA consists of two polynucleotide strands coiled around each other in a **double helix**. The two strands run in opposite directions and are held together by hydrogen bonds between specific pairs of bases. Adenine (A) and thymine (T) form strong hydrogen bonds to each other but not to C or G. Similarly, guanine (G) and cytosine (C) form strong hydrogen bonds to each other but not to A or T.



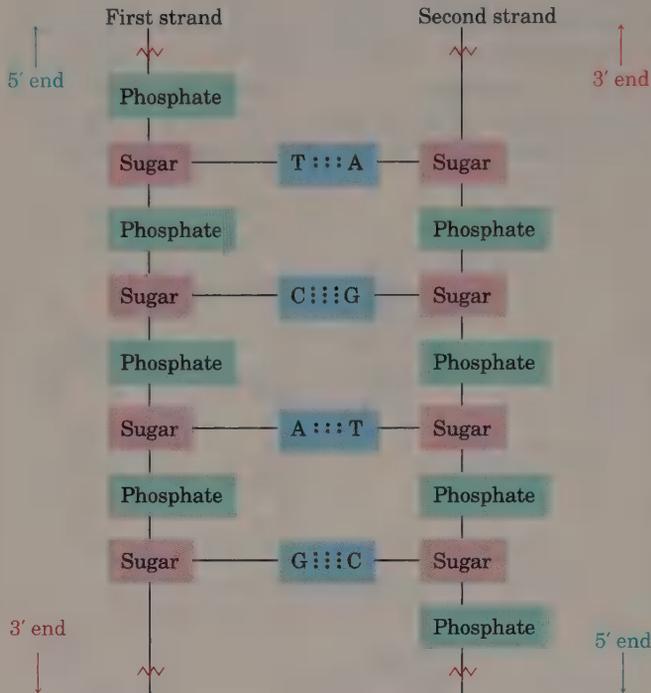
The two strands of the DNA double helix aren't identical; rather, they're complementary. Whenever a C base occurs in one strand, a G base occurs opposite it in the other strand. When an A base occurs in one strand, a T appears opposite it in the other strand. This complementary pairing of bases explains why A and T, and C and G, are always found in equal amounts. Figure 29.7 illustrates this base pairing, showing how the two complementary strands are coiled into the double helix. X-ray measurements show that the DNA double helix is 20 Å wide, that there are exactly 10 base pairs in each full turn, and that each turn is 34 Å in height.

A helpful mnemonic device to remember the nature of the hydrogen bonding between the four DNA bases is the simple phrase “Pure silver taxi.”

Pure	Silver	Taxi
Pur	Ag	TC
The purine bases,	A and G,	hydrogen-bond to T and C.

<sup>1</sup>James Dewey Watson (1928– ); b. Chicago, Ill.; Ph.D. Indiana; professor, Harvard University; Nobel Prize in medicine (1960).

<sup>2</sup>Francis H. C. Crick (1916– ); b. England; Ph.D. Cambridge; professor, Cambridge University; Nobel Prize in medicine (1960).



**Figure 29.7** Complementarity in base pairing in the DNA double helix as shown by this computer-generated structure. The sugar-phosphate backbone runs along the outside of the helix, while the amine bases hydrogen-bond to one another on the inside.

Notice in Figure 29.7 that the two strands of the double helix coil such that two kinds of “grooves” result, a **major groove** that’s 12 Å wide and a **minor groove** that’s 6 Å wide. The major groove is slightly deeper than the minor groove, and both are lined by potential hydrogen-bond donors and acceptors. Thus, a variety of molecules are able to *intercalate*, or fit into one of the grooves between the strands. A large number of cancer-causing and cancer-preventing agents are thought to function by interacting with DNA in this way.

**PROBLEM** .....

**29.13** What sequence of bases on one strand of DNA is complementary to the following sequence on another strand?

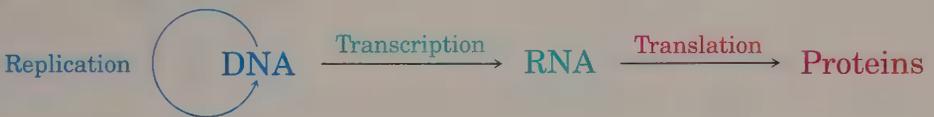
G-G-C-T-A-A-T-C-C-G-T

.....

## 29.12 Nucleic Acids and Heredity

A DNA molecule is the chemical repository of an organism's genetic information, which is stored as a sequence of deoxyribonucleotides strung together in the DNA chain. For this information to be preserved and passed on to future generations, a mechanism must exist for copying DNA. For the information to be used, a mechanism must exist for decoding the DNA message and for implementing the instructions it contains.

What Crick has termed the central dogma of molecular genetics says that the function of DNA is to store information and pass it on to RNA at the proper time. The function of RNA, in turn, is to read, decode, and use the information received from DNA to make proteins. Each of the thousands of individual genes on a chromosome contains the instructions necessary to make a specific protein needed for a specific biological purpose. By decoding the right genes at the right time in the right place, an organism can use genetic information to synthesize the many thousands of proteins necessary for carrying out the biochemical reactions required for smooth functioning.



Three fundamental processes take place in the transfer of genetic information:

1. **Replication** is the process by which identical copies of DNA are made so that information can be preserved and handed down to offspring.
2. **Transcription** is the process by which the genetic messages contained in DNA are read and carried out of the nucleus to parts of the cell called ribosomes where protein synthesis occurs.
3. **Translation** is the process by which the genetic messages are decoded and used to build proteins.

## 29.13 Replication of DNA

The Watson–Crick model of DNA does more than just explain base pairing; it also provides a remarkably ingenious way for DNA molecules to reproduce exact copies of themselves. **Replication** of DNA is an enzyme-catalyzed process that begins by a partial unwinding of the double helix. As the strands separate and bases are exposed, new nucleotides line up on each strand in an exactly complementary manner, A to T and C to G, and two new strands begin to grow. Each new strand is complementary to its old template strand, and two new identical DNA double helices are produced (Figure 29.8).

Crick described the process best when he used the analogy of the two DNA strands fitting together like a hand in a glove. The hand and glove

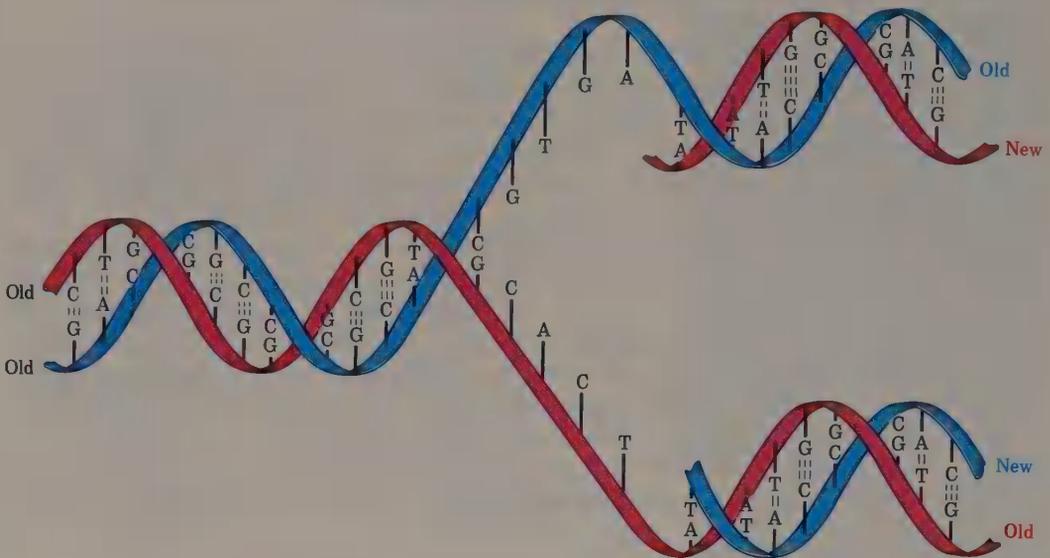


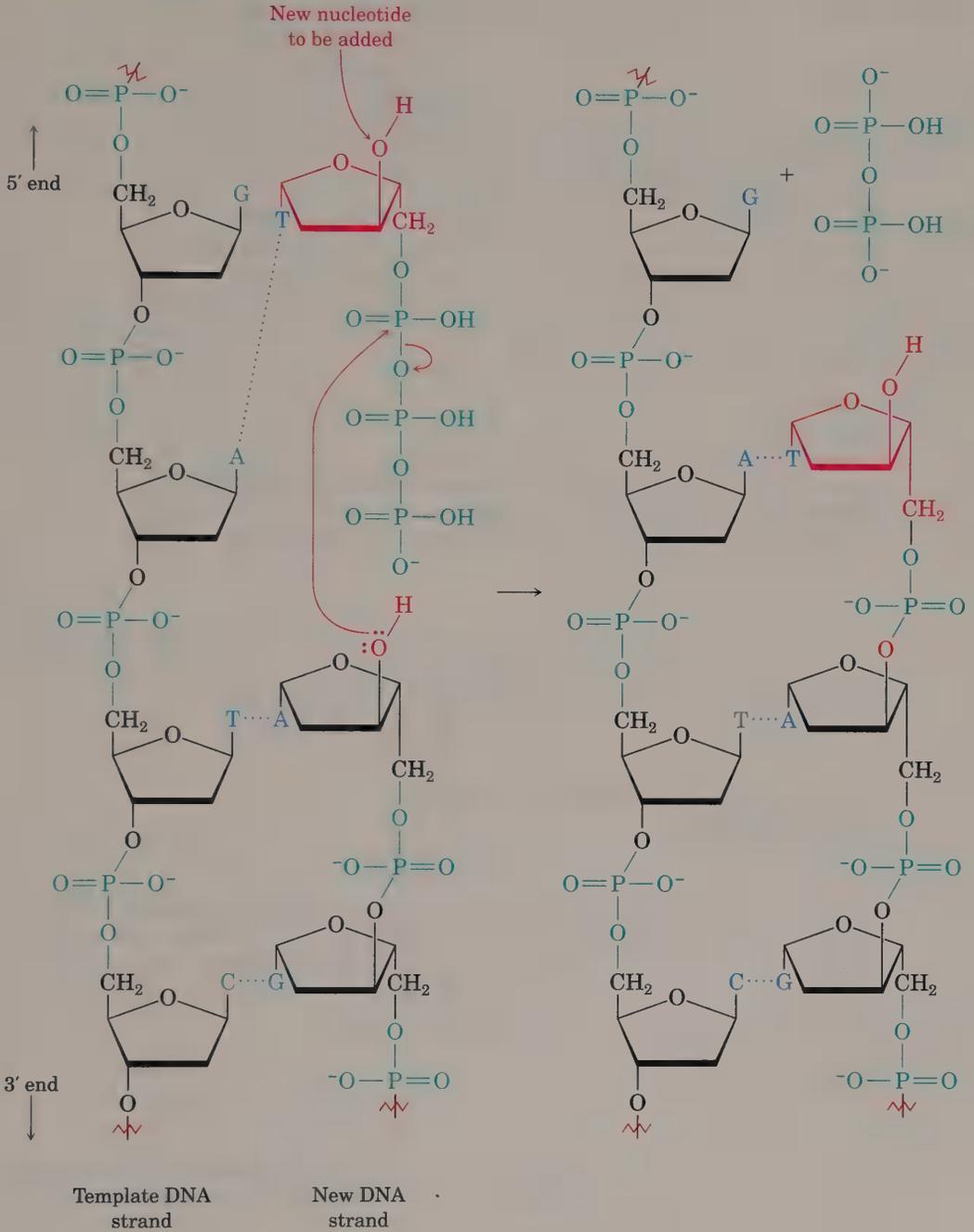
Figure 29.8 Schematic representation of DNA replication.

separate; a new hand forms inside the glove; and a new glove forms around the hand. Two identical copies now exist where only one existed before.

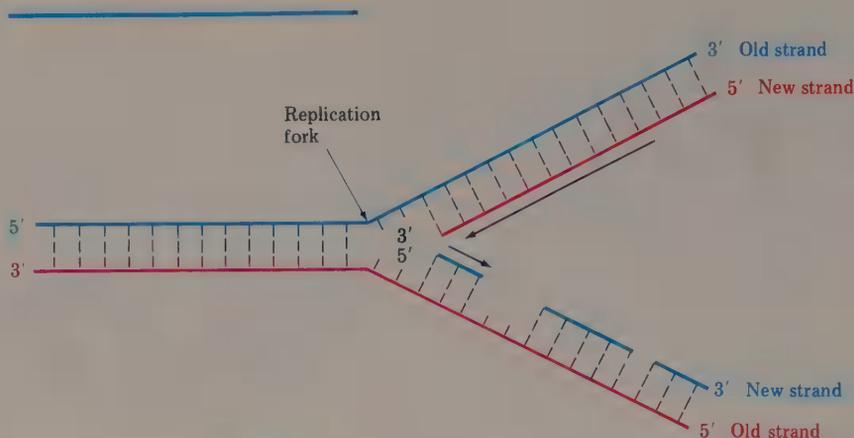
The process by which the individual nucleotides are joined to create new DNA strands involves many steps and many different enzymes. Addition of new nucleotide units to the growing chain is catalyzed by the enzyme *DNA polymerase* and has been shown to occur by addition of a 5'-mono-nucleotide triphosphate to the free 3'-hydroxyl group of the growing chain, as indicated in Figure 29.9.

Both of the new DNA strands are synthesized in the same 5'-to-3' direction, which implies that the two strands can't be synthesized in exactly the same way. Since the two complementary DNA strands are lined up in opposite directions, one strand must have its 3' end near the point of unraveling (the **replication fork**) while the other strand has its 5' end near the replication fork. What evidently happens is that the complement of the original 3' → 5' strand is synthesized smoothly and in a single piece, but the complement of the original 5' → 3' strand is synthesized discontinuously in small pieces that are then linked at a later point by DNA ligase enzymes (Figure 29.10).

It's difficult to conceive of the magnitude of the replication process. The nucleus of a human cell contains 46 chromosomes (23 pairs), each of which consists of one very large DNA molecule. Each chromosome, in turn, is made up of several thousand DNA segments called *genes*, and the sum of all genes in a human cell (the *genome*) is estimated to be approximately 3 billion base pairs. A single DNA chain might have a length of over 12 cm and contain up to 250 million pairs of bases. Regardless of the size of these enormous



**Figure 29.9** Addition of a new nucleotide to a growing DNA strand.

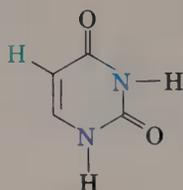


**Figure 29.10** Replication of DNA. Both new DNA strands are synthesized in a  $5' \rightarrow 3'$  direction. The original strand whose  $3'$  end is near the point of unraveling is replicated smoothly, but the original strand whose  $5'$  end is near the unraveling point is replicated in small pieces that are later joined. The arrowheads are at the growing ends of the chains.

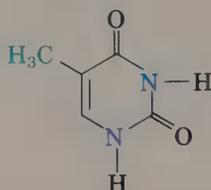
molecules, their base sequence is faithfully copied during replication. The copying process takes only minutes, and it's been estimated that an error occurs only about once each 10 billion to 100 billion bases.

## 29.14 Structure and Synthesis of RNA: Transcription

RNA is structurally similar to DNA. Both are sugar–phosphate polymers and both have heterocyclic bases attached. The only differences are that RNA contains ribose rather than deoxyribose and uracil rather than thymine. Uracil in RNA forms strong hydrogen bonds to its complementary base, adenine, just as thymine does in DNA.



Uracil (in RNA)



Thymine (in DNA)

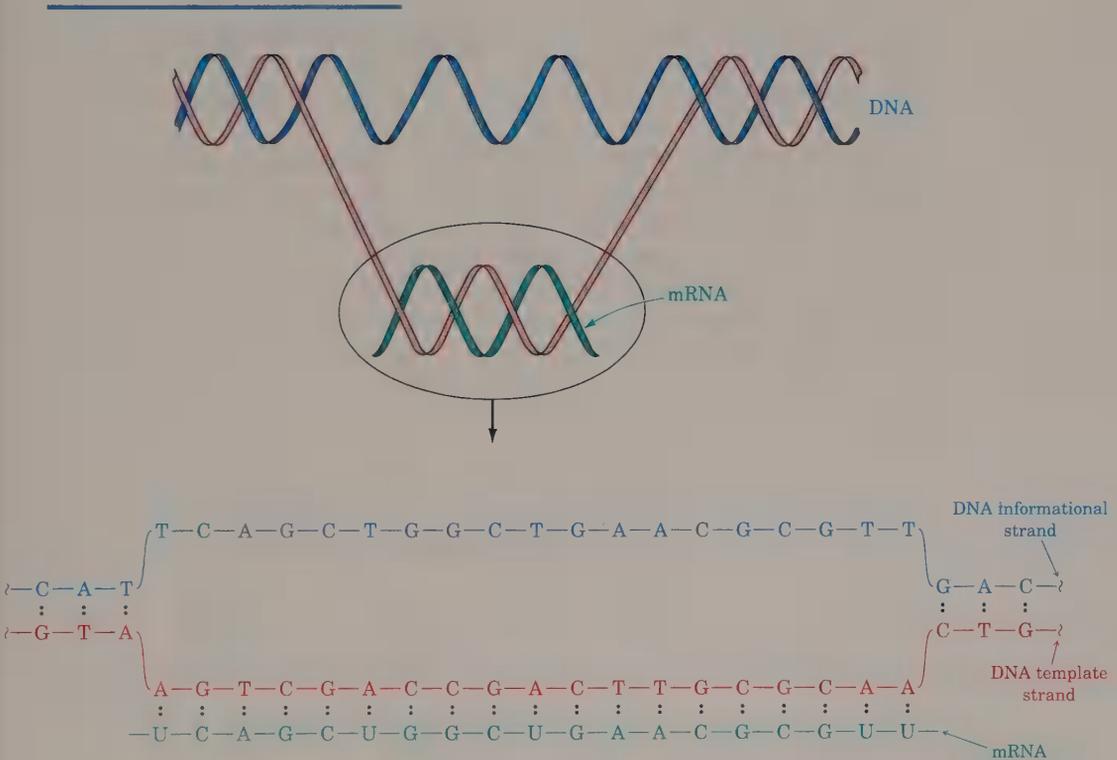
There are three major kinds of ribonucleic acid, each of which serves a specific function. All three kinds of RNA are structurally similar to DNA, but they're much smaller molecules than DNA and they remain single stranded, rather than double stranded like DNA.

**Messenger RNA (mRNA)** carries genetic messages from DNA to *ribosomes*, small granular particles in the cytoplasm of a cell where protein synthesis takes place.

**Ribosomal RNA (rRNA)** complexed with protein provides the physical makeup of the ribosomes.

**Transfer RNA (tRNA)** transports specific amino acids to the ribosomes where they're joined together to make proteins.

Molecules of RNA are synthesized in the nucleus of the cell by **transcription** of DNA. A portion of the DNA double helix unwinds, and the bases of the two strands are exposed. Ribonucleotides line up in the proper order by hydrogen-bonding to their complementary bases on DNA, bond formation occurs in the 5' → 3' direction, and the completed RNA molecule then unwinds from DNA and migrates from the nucleus (Figure 29.11).



**Figure 29.11** Biosynthesis of RNA using a DNA segment as template.

The conversion of the information contained in a DNA segment into proteins begins with the synthesis of mRNA molecules that contain anywhere from several hundred to several thousand ribonucleotides, depending on the size of the protein to be made. Each of the 100,000 or so proteins in the human body is synthesized from a different mRNA that has been transcribed from a specific gene on DNA.

Unlike what happens in DNA replication, where both strands are copied, only one of the two DNA strands is transcribed into mRNA. The strand that contains the gene is called the **informational strand**, and its complement is called the **template strand**. Only the template strand is transcribed. Since the template strand and the informational strand are complementary, and since the template strand and the mRNA molecule are also complementary, it follows that *the messenger RNA molecule produced during transcription is a copy of the informational strand*. The only difference is that the mRNA molecule has a uracil everywhere the DNA informational strand has a thymine.

Transcription of DNA by the process just discussed raises as many questions as it answers. How does the DNA know where to unwind? Where along the chain does one gene stop and the next one start? How do the ribonucleotides know the right place along the template strand to begin lining up and the right place to stop? Though these questions are extremely difficult to answer, the picture that has emerged in the last decade is that a DNA chain contains certain base sequences called **promoter sites** that bind the RNA polymerase enzyme that actually carries out RNA synthesis, thus signaling the beginning of a gene. Similarly there are other base sequences at the end of the gene that signal a stop.

Another part of the picture that has recently emerged is that genes are not necessarily continuous segments of the DNA chain. Often a gene will begin in one small section of DNA called an **exon**, then be interrupted by a seemingly nonsensical section called an **intron**, and then take up again further down the chain in another exon. The final mRNA molecule results only after the nonsense sections are cut out and the remaining pieces are spliced together. Current evidence is that up to 90% of human DNA is made up of introns and only about 10% of DNA actually contains genetic instructions.

PROBLEM .....

- 29.14 Show how uracil can form strong hydrogen bonds to adenine.

PROBLEM .....

- 29.15 What RNA base sequence is complementary to the following DNA base sequence?

G-A-T-T-A-C-C-G-T-A

.....

## 29.15 RNA and Protein Biosynthesis: Translation

---

The primary cellular function of RNA is to direct biosynthesis of the thousands of diverse peptides and proteins required by an organism. These proteins in turn regulate all other biological processes. The mechanics of protein biosynthesis are directed by messenger RNA (mRNA) and take place on ribosomes, small granular particles in the cytoplasm of a cell that consist

of about 60% ribosomal RNA and 40% protein. On the ribosome, mRNA serves as a template to pass on the genetic information it has transcribed from DNA.

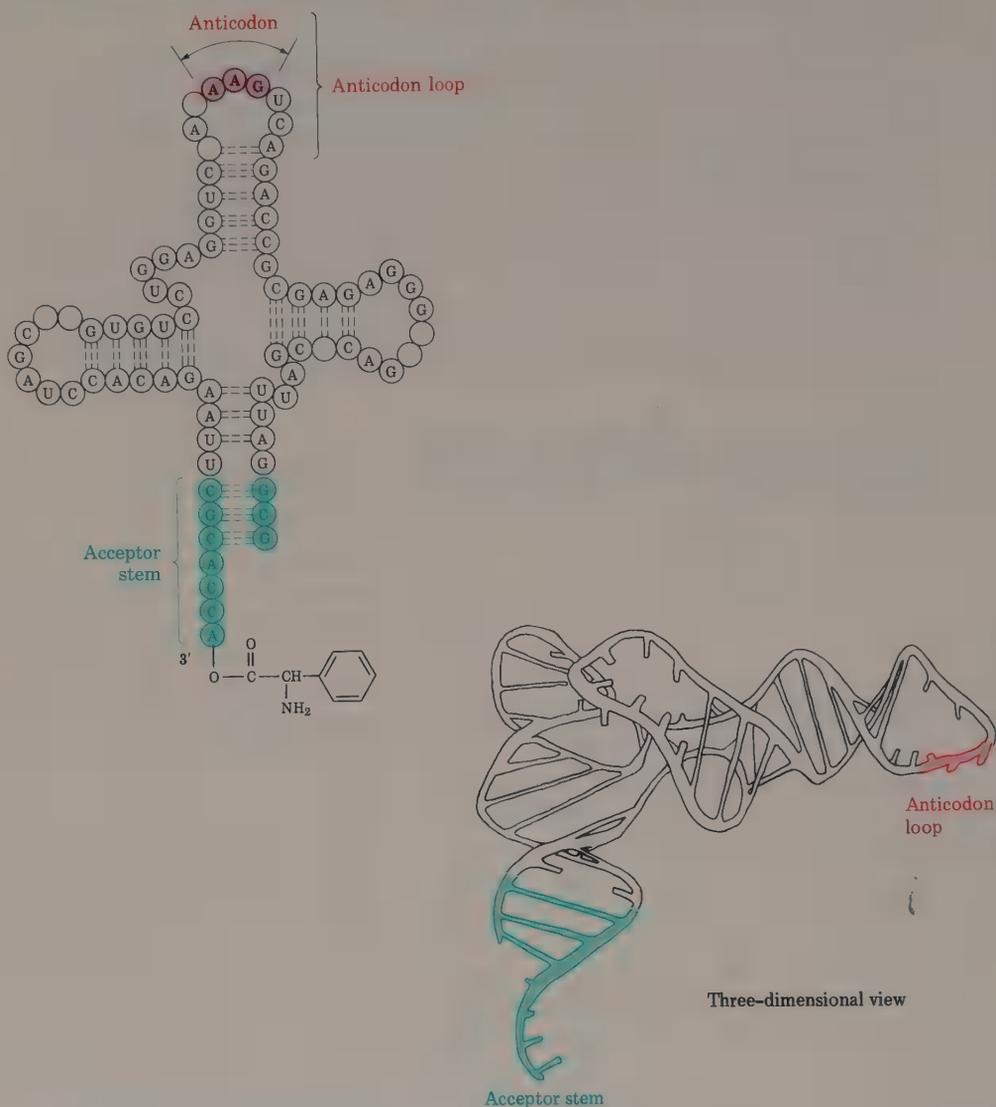
The specific ribonucleotide sequence in mRNA forms a “code” that determines the order in which different amino acid residues are to be joined. Each “word” or **codon** along the mRNA chain consists of a series of three ribonucleotides that is specific for a given amino acid. For example, the series cytosine–uracil–guanine (C-U-G) on mRNA is a codon directing incorporation of the amino acid leucine into the growing protein, and guanine–adenine–uracil (G-A-U) codes for aspartic acid. Of the  $4^3 = 64$  possible triads of the four bases in RNA, 61 code for specific amino acids (certain amino acids are specified by more than one codon). In addition, 3 of the 64 codons specify chain termination. Table 29.3 shows the meaning of each codon.

**TABLE 29.3** Codon Assignments of Base Triads

First base (5' end)	Second base	Third base (3' end)			
		U	C	A	G
U	U	Phe	Phe	Leu	Leu
	C	Ser	Ser	Ser	Ser
	A	Tyr	Tyr	Stop	Stop
	G	Cys	Cys	Stop	Trp
C	U	Leu	Leu	Leu	Leu
	C	Pro	Pro	Pro	Pro
	A	His	His	Gln	Gln
	G	Arg	Arg	Arg	Arg
A	U	Ile	Ile	Ile	Met
	C	Thr	Thr	Thr	Thr
	A	Asn	Asn	Lys	Lys
	G	Ser	Ser	Arg	Arg
G	U	Val	Val	Val	Val
	C	Ala	Ala	Ala	Ala
	A	Asp	Asp	Glu	Glu
	G	Gly	Gly	Gly	Gly

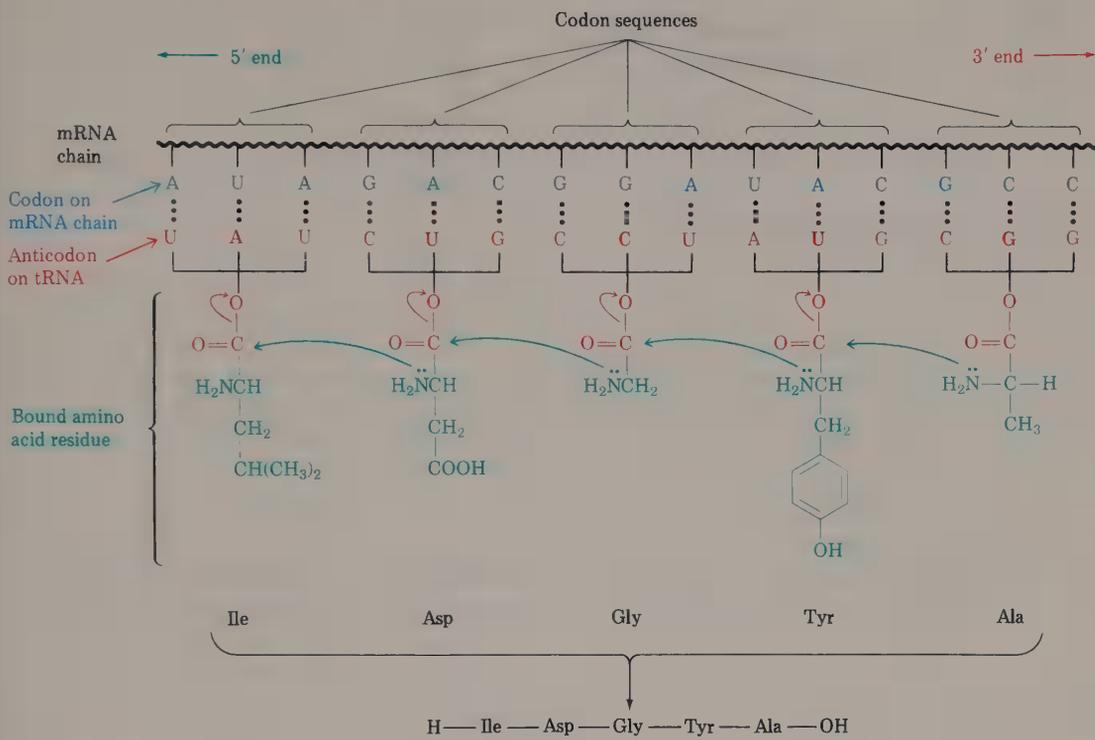
The code expressed in mRNA is read by transfer RNA (tRNA) in a process called **translation**. There are at least 60 different transfer RNA's, one for each of the codons in Table 29.3. Each specific tRNA acts as a carrier to bring a specific amino acid into place so that it can be transferred to the growing protein chain.

A typical tRNA is roughly the shape of a cloverleaf, as shown in Figure 29.12. It consists of about 70–100 ribonucleotides and is bonded to a specific amino acid by an ester linkage through the free 3'-hydroxyl on ribose at the 3' end of the tRNA. Each tRNA also contains in its structure a segment called an **anticodon**, a sequence of three ribonucleotides complementary to the codon sequence. For example, the codon sequence C-U-G present on mRNA is read by a leucine-bearing tRNA having the complementary anticodon base sequence G-A-C.



**Figure 29.12** Structure of a tRNA molecule. The tRNA is a roughly cloverleaf-shaped molecule containing an anticodon triplet on one “leaf” and a covalently attached amino acid unit at its 3’ end. The example shown is a yeast tRNA that codes for phenylalanine. The nucleotides not specifically identified are chemically modified analogs of the four normal nucleotides.

As each successive codon on mRNA is read, different tRNA’s bring the correct amino acids into position for enzyme-mediated transfer to the growing peptide. When synthesis of the proper protein is completed, a “stop” codon signals the end, and the protein is released from the ribosome. The entire process of protein biosynthesis is illustrated schematically in Figure 29.13.



**Figure 29.13** A schematic representation of protein biosynthesis. The mRNA containing codon base sequences is read by tRNA containing complementary anticodon base sequences. Transfer RNA assembles the proper amino acids into position for incorporation into the growing peptide.

**PROBLEM** .....

- 29.16** List codon sequences for these amino acids:  
 (a) Ala                      (b) Phe                      (c) Leu                      (d) Tyr

**PROBLEM** .....

- 29.17** List anticodon sequences on the tRNA's carrying the amino acids shown in Problem 29.16.

**PROBLEM** .....

- 29.18** What amino acid sequence is coded for by the following mRNA base sequence?

CUU-AUG-GCU-UGG-CCC-UAA

**PROBLEM** .....

- 29.19** What anticodon sequences of tRNA's are coded for by the mRNA in Problem 29.18?

## PROBLEM .....

- 29.20 What is the base sequence in the original DNA strand on which the mRNA sequence in Problem 29.18 was made?
- .....

## 29.16 Sequencing of DNA

---

When we work out the structure of DNA molecules, we examine the fundamental level that underlies all processes in living cells. DNA is the information store that ultimately dictates the structure of every gene product, delineates every part of the organism. The order of the bases along DNA contains the complete set of instructions that make up the genetic inheritance. (Walter Gilbert, Nobel Prize Lecture, 1980)

One of the greatest scientific revolutions in history is now under way in molecular biology as scientists are learning how to manipulate and harness the genetic machinery of organisms. None of the extraordinary advances of the past decade would have been possible, however, were it not for the discovery in 1977 of a method for sequencing immense DNA chains to find the messages therein. Much DNA sequencing is now carried out by a remarkably efficient and powerful method developed by Allan Maxam and Walter Gilbert.<sup>3</sup> There are five steps:

**Step 1** Since molecules of DNA are so enormous—some molecules of human DNA contain up to 250 million base pairs—the first problem in DNA sequencing is to find a method for cleaving the DNA chain at specific points to produce smaller, more manageable pieces. This problem has been solved by the use of enzymes called **restriction endonucleases**. Each different restriction enzyme, of which more than 200 are available, cleaves a DNA molecule between two nucleotides at well-defined points along the chain where specific base sequences occur. Since the required sequence is usually four or more nucleotides long, it is unlikely to occur very often in the overall DNA sequence.

By cleavage of a large DNA molecule with a given restriction enzyme, many different and well-defined segments of manageable length (100–200 nucleotides) are produced. For example, the restriction enzyme *Alu I* cleaves the linkage between G and C in the four-base sequence AG-CT, and the enzyme *Hpa II* cleaves the C-C linkage in the four-base sequence C-CGG (Figure 29.14). If the original DNA molecule is cut with another restriction enzyme having a different specificity for cleavage, still other segments are produced whose sequences partially overlap those produced by the first enzyme. Sequencing of all the segments, followed by identification of the overlapping sequences, then allows complete DNA sequencing in a manner similar to that used for protein sequencing (Section 27.9).

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<sup>3</sup>Walter Gilbert (1932– ); b. Boston; Ph.D. Cambridge University (1957); professor, Harvard University (1958– ); Nobel Prize (1980).

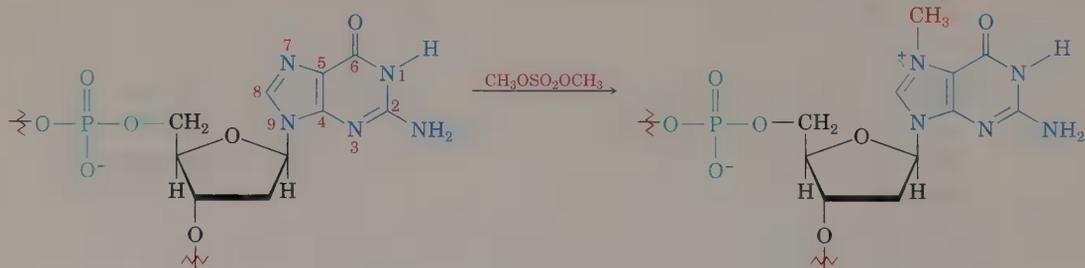


Mild reaction conditions are chosen so that *only a few of the many possible splittings occur in each reaction*. In our example, the pieces shown in Table 29.4 would be produced.

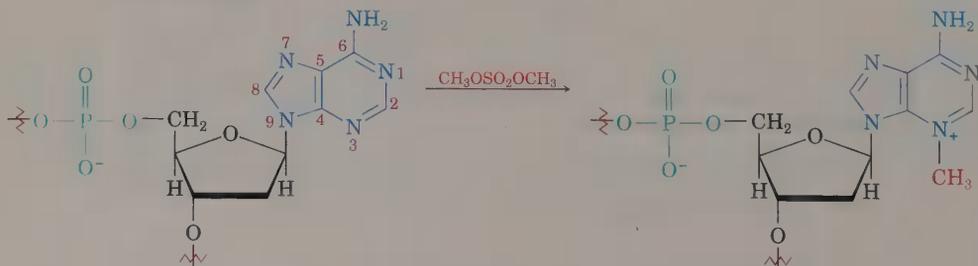
**TABLE 29.4** Splitting of a DNA Fragment Under Four Sets of Conditions

Cleavage conditions	Labeled DNA pieces produced
Original DNA fragment	$^{32}\text{P}$ -G-A-T-C-A-G-C-G-A-T-
next to A	$^{32}\text{P}$ -G $^{32}\text{P}$ -G-A-T-C $^{32}\text{P}$ -G-A-T-C-A-G-C-G + larger pieces
next to G	$^{32}\text{P}$ -G-A-T-C-A $^{32}\text{P}$ -G-A-T-C-A-G-C + larger pieces
next to C	$^{32}\text{P}$ -G-A-T $^{32}\text{P}$ -G-A-T-C-A-G + larger pieces
next to C + T	$^{32}\text{P}$ -G-A $^{32}\text{P}$ -G-A-T $^{32}\text{P}$ -G-A-T-C-A-G $^{32}\text{P}$ -G-A-T-C-A-G-C-G-A + larger pieces

Cleavages next to A and G are accomplished by treatment of a restriction fragment with dimethyl sulfate  $[(\text{CH}_3\text{O})_2\text{SO}_2]$ . Deoxyadenosine (A) is methylated at N3 ( $\text{S}_{\text{N}}2$  reaction), and deoxyguanosine (G) is methylated at N7, but T and C aren't affected.



**Deoxyguanosine**



**Deoxyadenosine**

Treatment of the methylated DNA with an aqueous solution of the secondary amine piperidine then brings about destruction of the methylated nucleotides and opening of the DNA chain at both the 3' and 5' positions next to the methylated bases. The mechanism of the cleavage process is complex and involves a number of different steps, as shown in Figure 29.15 for deoxyguanosine breakage:

1. Hydrolysis occurs, opening the five-membered heterocycle.
2. Cleavage of the aminoglycoside sugar linkage then yields an open-chain 2-deoxyribose.
3. Formation of an enamine between piperidine and the 2-deoxyribose aldehyde group occurs.
4. Two eliminations of the 2-deoxyribose oxygen substituents at C3 and C5 take place. It's these two elimination reactions that break open the DNA chain.

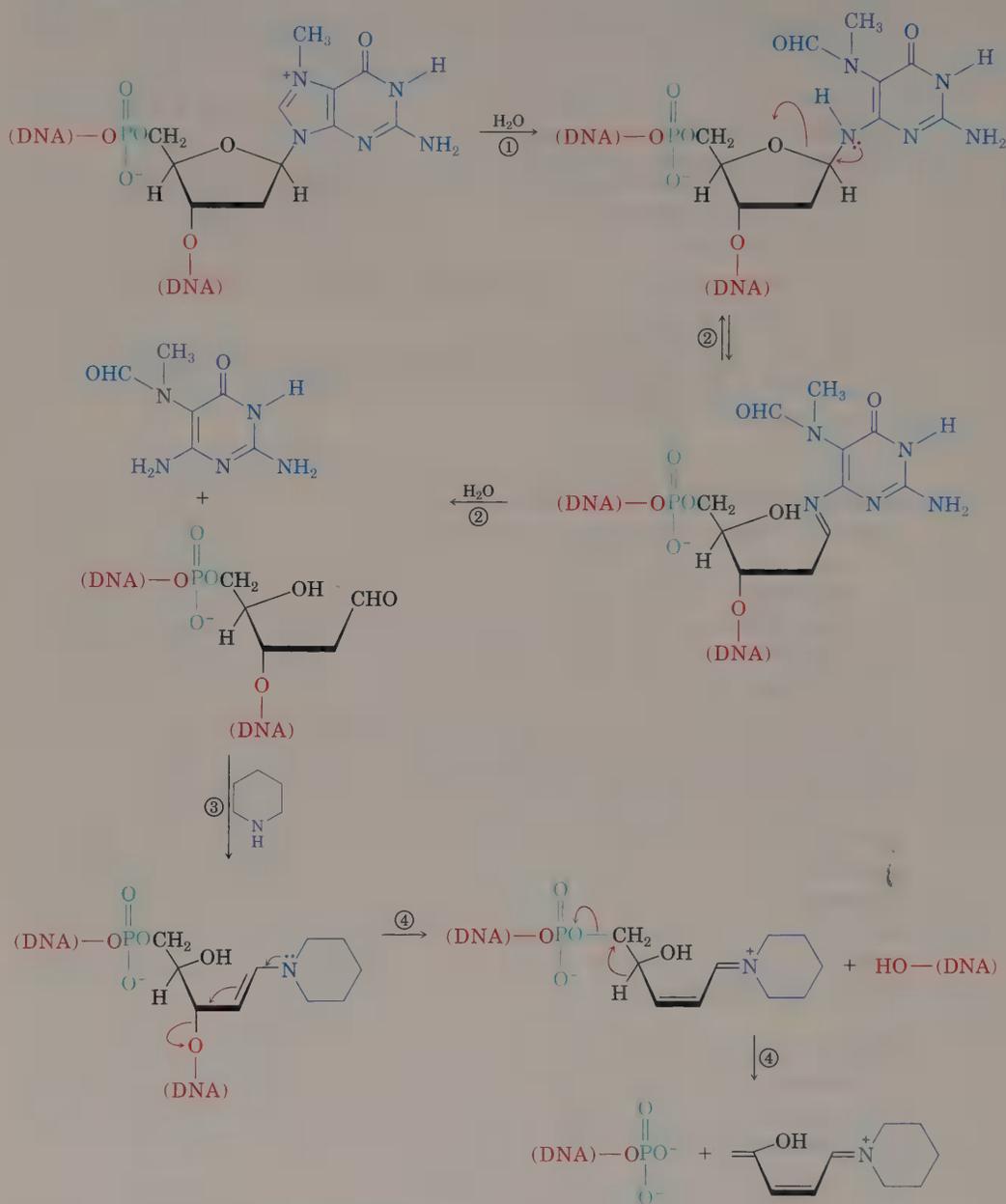
By working carefully, Maxam and Gilbert were able to find reaction conditions that are selective for cleavage either at A or at G. Thus, G methylates five times as rapidly as A but the hydrolytic breakdown of methylated A occurs more rapidly than the corresponding breakdown of methylated G if the product is first heated with dilute formic acid prior to base treatment.

Breaking the DNA chain next to both pyrimidine nucleotides C and T is accomplished by treatment of DNA with hydrazine, followed by heating with aqueous piperidine. Although no conditions have been found that are selective for cleavage next to T, a selective cleavage next to C can be accomplished by carrying out the hydrazine reaction in 5 M NaCl. The mechanism of deoxythymidine cleavage by hydrazine is shown in Figure 29.16.

Once again, the breakdown reaction involves numerous steps: (1) Hydrazine first undergoes a conjugate addition to thymine, followed by (2) an intramolecular nucleophilic acyl substitution reaction that opens the heterocyclic thymine ring; (3) breakage of the aminoglycoside linkage occurs, and (4) hydrolysis yields open-chain deoxyribose, which (5) forms an enamine by reaction with piperidine; (6) two elimination reactions then break open the DNA chain.

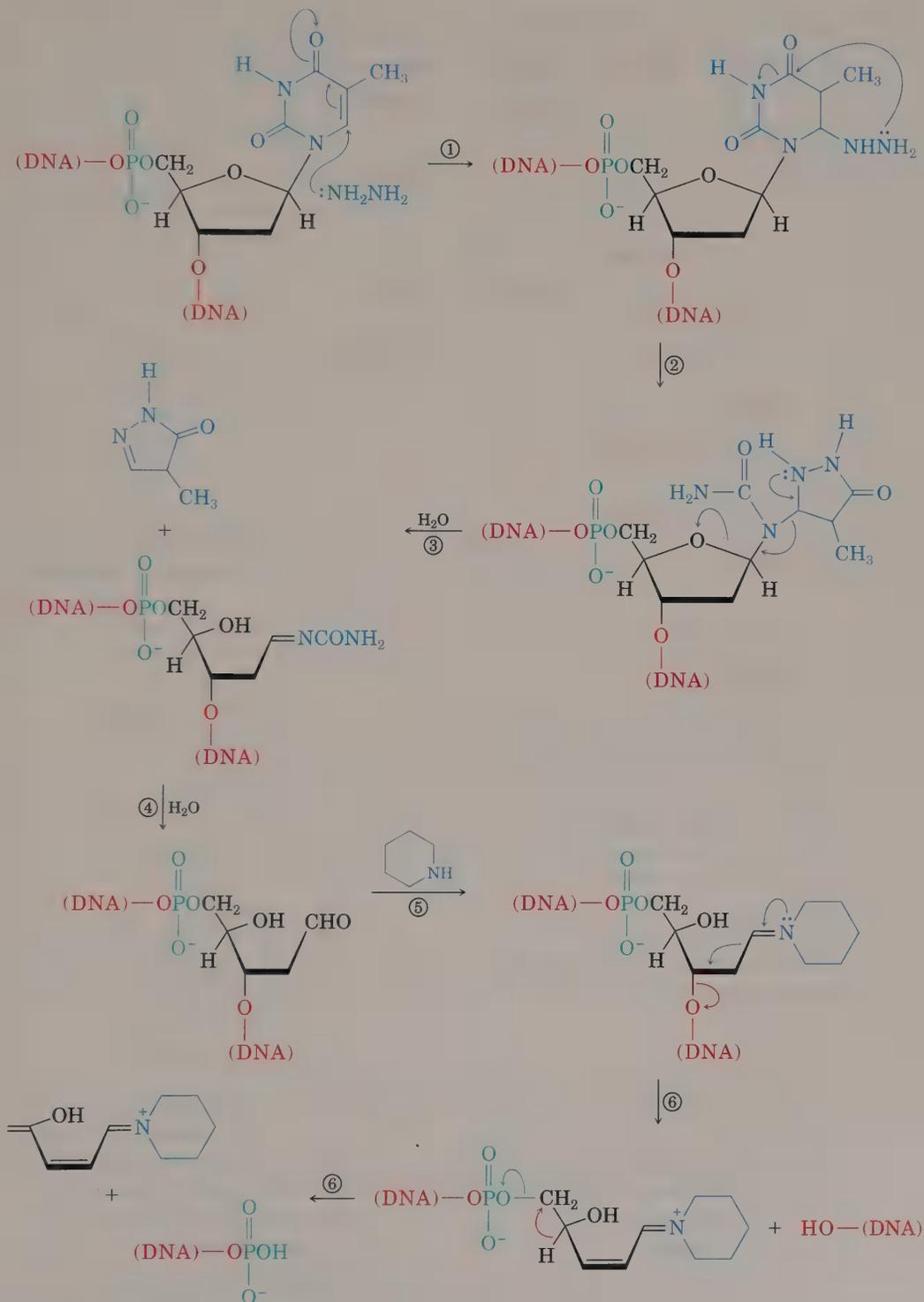
**Step 4** Product mixtures from the four cleavage reactions are separated by electrophoresis (Section 27.3). When the mixture of DNA pieces is placed at one end of a strip of buffered gelatinous polyacrylamide and a voltage difference is applied across the ends of the strip, electrically charged pieces move along the gel. Each piece moves at a rate that depends both on its size and on the number of negatively charged phosphate groups (the number of nucleotides) it contains. Smaller pieces move rapidly, and larger pieces move more slowly. The technique is so sensitive that up to 600 DNA pieces, differing in size by only one nucleotide, can be separated.

Once separated, the locations of each DNA fragment are detected by exposing the gel to a photographic plate. Each radioactive end piece containing a  $^{32}\text{P}$  label appears as a dark band on the photographic plate, but nonradioactive pieces from the middle of the chain aren't visualized. The gel electrophoresis pattern shown in Figure 29.17 would be obtained in our hypothetical example.

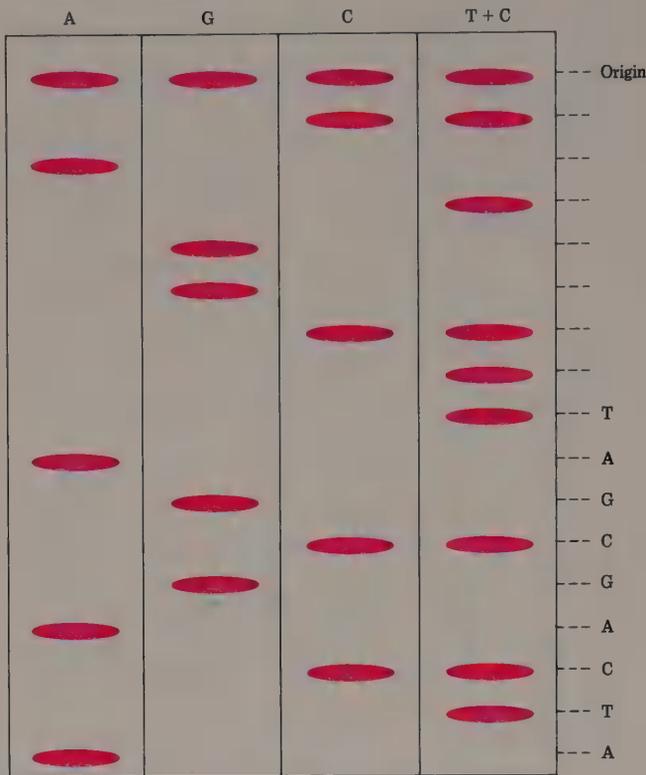


**Figure 29.15** Mechanism of DNA cleavage at deoxyguanosine (G).

**Step 5** The DNA sequence is then read directly from the gel. The band that appears farthest from the origin is the terminal mononucleotide (the smallest piece) and can't be identified. Since the terminal mononucleotide appears in the A column, though, it must have been produced by splitting *next to* an A. Thus, the *second* nucleotide in the sequence is an A.



**Figure 29.16** Mechanism of DNA cleavage at deoxythymidine (T) by reaction with hydrazine.



**Figure 29.17** Representation of a gel electrophoresis pattern. The products of the four cleavage experiments are placed at the top of the gel, and a voltage is applied between top and bottom. Smaller products migrate along the gel at a faster rate and thus appear at the bottom. The DNA sequence can be read from the positions of the radioactive spots.

The second farthest band from the origin is a dinucleotide that appears only in the T + C column and is produced by splitting next to the third nucleotide, which must therefore be a T or a C. Since this piece doesn't appear in the C column, though, the third nucleotide isn't a C and must therefore be a T. The third farthest band appears in both C and T + C columns, meaning that the fourth nucleotide is a C.

Continuing in this manner, the entire sequence of the DNA can be read from the gel simply by noting in what columns the successively larger labeled polynucleotide pieces appear. Once read, the entire sequence can be checked by determining the sequence of the complementary strand. The identity of the 5'-terminal nucleotide can be determined by sequencing an overlapping segment produced by cleavage with another restriction enzyme.

The Maxam–Gilbert method of DNA sequencing is so efficient that a single worker can sequence up to 2000 base pairs per day. DNA strands of up to 170,000 base pairs have already been sequenced, and work is now under way to sequence the entire human genome with 3,000,000,000 base pairs. At least 10 years and several billion dollars will be needed.

## PROBLEM .....

- 29.21 Show the labeled cleavage products you would expect to obtain if the following DNA segment were subjected to each of the four cleavage reactions:



## PROBLEM .....

- 29.22 Sketch what you would expect the gel electrophoresis pattern to look like if the DNA segment in Problem 29.21 were sequenced.

## PROBLEM .....

- 29.23 Finish assigning the sequence to the gel electrophoresis pattern shown in Figure 29.17.
- .....

## 29.17 Laboratory Synthesis of DNA

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The development of genetic engineering techniques in the last two decades brought with it an increased demand for efficient chemical methods for the synthesis of short DNA segments. Ideally, whole genes might be synthesized in the laboratory and inserted into the DNA of microorganisms, thereby directing the microorganisms to produce the specific protein coded for by that gene—perhaps insulin or some other valuable material.

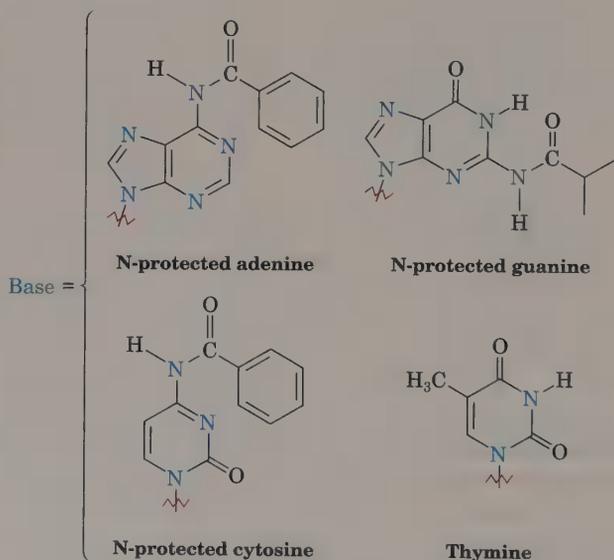
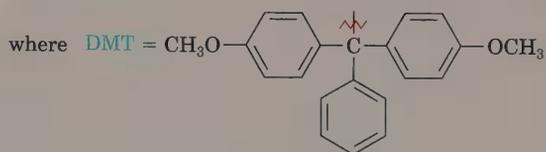
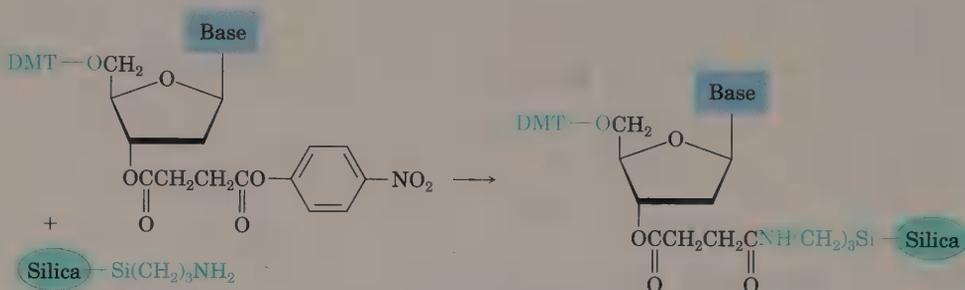
The problems of DNA synthesis are similar to those of protein synthesis (Section 27.11) but are considerably more difficult because of the structural complexity of the deoxyribonucleotide monomers. Each nucleotide has several reactive sites that must be selectively protected and deprotected at the proper times, and coupling of the four nucleotides must be carried out in the proper sequence. Despite these difficulties, some extremely impressive early achievements were recorded, most notably the synthesis by Khorana<sup>4</sup> in 1979 of the tyrosine suppressor tRNA gene from the bacterium *Escherichia coli*. Some 207 base pairs were assembled in an effort that required 10 years of work.

More recently, automated “gene machines” have become available, which allow the fast and reliable synthesis of DNA sequences 100–200 nucleotides long. These DNA synthesizers operate on a principle similar to that of the Merrifield solid-phase peptide synthesizer (Section 27.12). In essence, a protected nucleotide is covalently bound to a solid support, and one nucleotide at a time is added to the chain. When the last nucleotide has been added, the protecting groups are removed, and the synthetic DNA is cleaved from the solid support.

<sup>4</sup>Har Gobind Khorana (1922– ); b. Raipur, India; Ph.D. University of Liverpool; professor, Massachusetts Institute of Technology; Nobel Prize in medicine (1968).

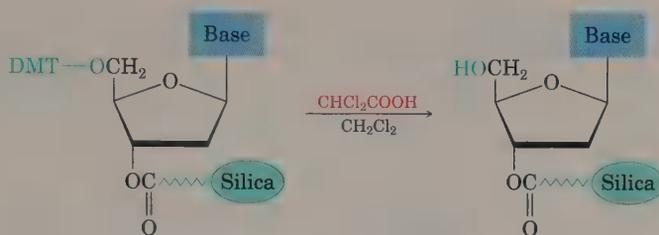
Step 1 in DNA synthesis involves attachment of a protected deoxynucleoside to a silica ( $\text{SiO}_2$ ) support by an ester linkage to the deoxynucleoside's 3' position. Both the 5' hydroxyl and free amino groups on the heterocyclic bases must be protected to accomplish this attachment. Adenine and cytosine bases are protected by benzoyl groups, guanine is protected by an isobutyryl group, and thymine requires no protection. The deoxyribose 5' hydroxyl is protected as its DMT ether, where DMT = *p*-dimethoxytrityl. Other protecting groups can also be used, but those mentioned here are common choices.

Step 1



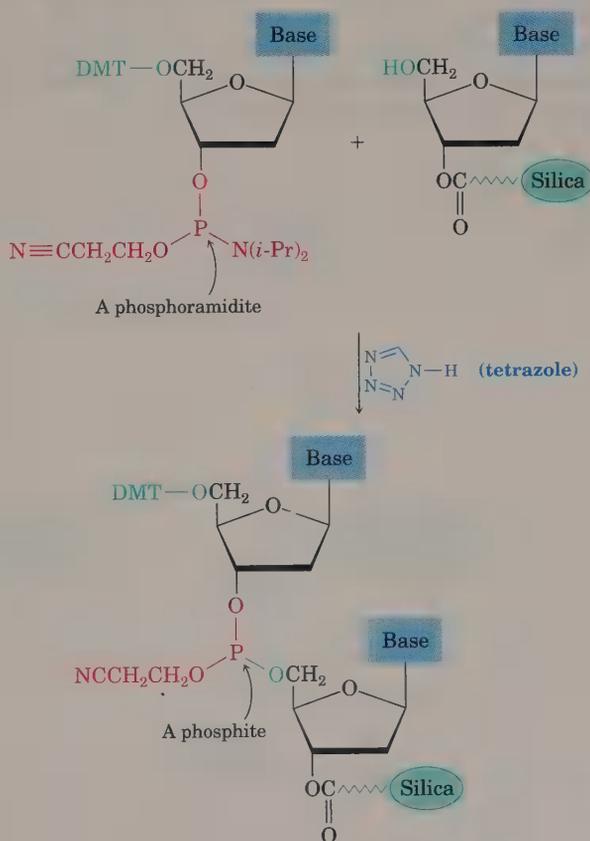
Step 2 involves deprotection of the 5'-deoxyribose hydroxyl by treatment with dichloroacetic acid in  $\text{CH}_2\text{Cl}_2$  to remove the DMT group. The reaction proceeds rapidly by an  $\text{S}_{\text{N}}1$  mechanism because of the stability of the tertiary, triphenyl dimethoxytrityl cation.

## Step 2



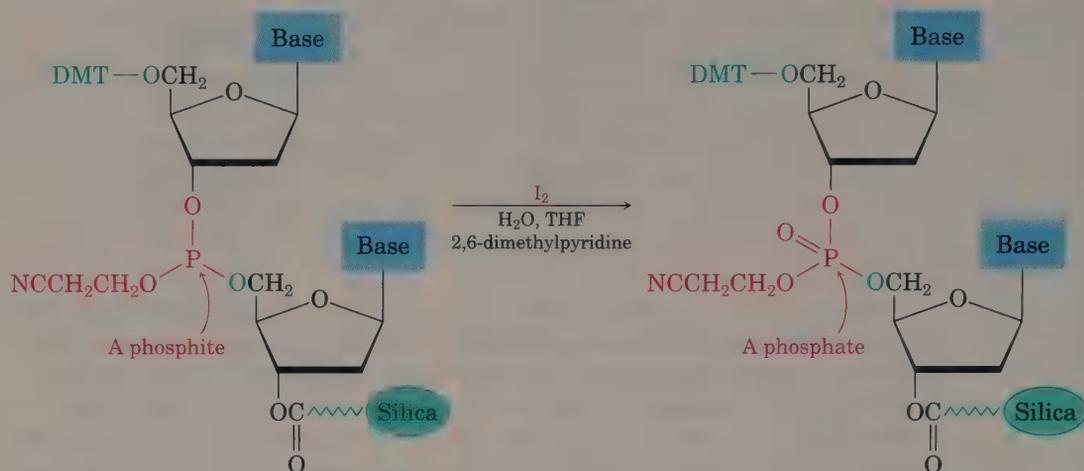
Step 3 involves coupling of the polymer-bonded deoxynucleoside with a protected deoxynucleoside containing a *phosphoramidite* group at its 3' position. [A phosphoramidite has a trivalent phosphorus with the structure  $\text{R}_2\text{NP}(\text{OR})_2$ .] The coupling reaction takes place in the polar aprotic solvent acetonitrile, requires catalysis by the heterocyclic amine tetrazole, and yields a *phosphite*, or trialkoxyphosphorus compound,  $\text{P}(\text{OR})_3$ , as product. Note that one of the phosphorus oxygen atoms is protected by a  $\beta$ -cyanoethyl group. The coupling step takes place in better than 99% yield.

## Step 3



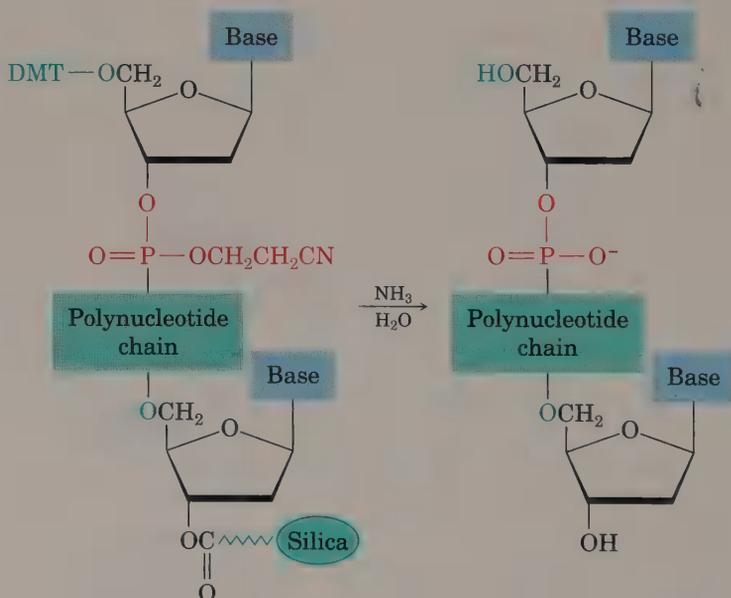
With the coupling completed, the phosphite product is then oxidized to a phosphate triester by treatment with iodine. The reaction is carried out in aqueous tetrahydrofuran in the presence of 2,6-dimethylpyridine.

## Step 4



The cycle of steps—deprotection, coupling, and oxidation—is then repeated until a polydeoxyribonucleotide chain of the desired length and sequence has been built. The final step is to cleave all protecting groups from the heterocyclic bases and from the phosphates and to cleave the ester bond holding the DNA to the silica. All of these reactions are done at the same time by treatment with aqueous ammonia. Purification by electrophoresis then yields the synthetic DNA.

## Step 5



## PROBLEM.....

- 29.24 *p*-Dimethoxytrityl (DMT) ethers are easily cleaved by mild acid treatment. Show the mechanism, and suggest a reason why this ether cleavage is unusually easy.

## PROBLEM.....

- 29.25 Propose a mechanism to account for cleavage of the  $\beta$ -cyanoethyl protecting group from the phosphate groups on treatment with aqueous ammonia. What kind of reaction is occurring?
- .....

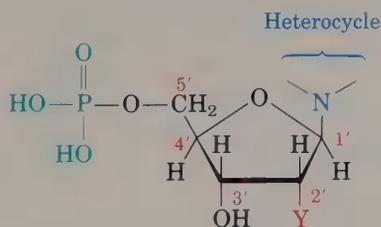
## 29.18 Summary and Key Words

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A **heterocycle** is a compound with a ring that contains more than one kind of atom. Nitrogen, oxygen, and sulfur are often found along with carbon in heterocyclic rings. Saturated heterocyclic amines, ethers, and sulfides usually display the same chemistry as their open-chain analogs do, but unsaturated heterocycles often display aromaticity. Pyrrole, furan, and thiophene are the simplest five-membered aromatic heterocycles. All three are unusually stable, and all three undergo aromatic substitution when reacted with electrophiles. Reaction usually occurs at the highly activated position next to the heteroatom.

Pyridine is the six-membered-ring, nitrogen-containing heterocyclic analog of benzene. The pyridine ring is electron-poor and undergoes electrophilic aromatic substitution reactions with great difficulty. Nucleophilic aromatic substitutions of 2- or 4-halopyridines take place readily, however.

The nucleic acids—DNA (**deoxyribonucleic acid**) and RNA (**ribonucleic acid**)—are biological polymers that act as chemical carriers of an organism's genetic information. Enzyme-catalyzed hydrolysis of nucleic acids yields **nucleotides**, the monomer units from which RNA and DNA are constructed. Each nucleotide consists of a **purine** or **pyrimidine base** linked to C1' of a simple pentose sugar (ribose in RNA and 2'-deoxyribose in DNA), with the sugar in turn linked through its C5' hydroxyl to a phosphate group:



where Y = OH (a ribonucleotide)  
 . = H (a deoxyribonucleotide)

The nucleotides are joined by phosphate links between the phosphate of one nucleotide and the 3'-hydroxyl on the sugar of another nucleotide.

Molecules of DNA consist of two polynucleotide strands held together by hydrogen bonds between heterocyclic bases on the different strands and coiled into a **double-helix** conformation. **Adenine** and **thymine** form hydrogen bonds to each other, as do **cytosine** and **guanine**. The two strands of DNA are not identical but are complementary.

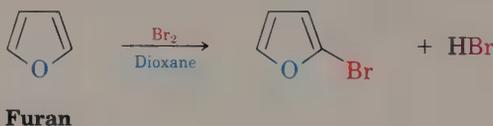
Three main processes take place in deciphering the genetic information of DNA:

1. **Replication** of DNA is the process by which identical DNA copies are made and genetic information is preserved. This occurs when the DNA double helix unwinds, complementary deoxyribonucleotides line up in order, and two new DNA molecules are produced.
2. **Transcription** is the process by which RNA is produced to carry genetic information from the nucleus to the ribosomes. This occurs when a segment of the DNA double helix unwinds and complementary ribonucleotides line up to produce **messenger RNA (mRNA)**.
3. **Translation** is the process by which mRNA directs protein synthesis. Each mRNA has segments called **codons** along its chain. These codons are ribonucleotide triads that are recognized by small amino acid-carrying molecules of **transfer RNA (tRNA)**, which then deliver the appropriate amino acids needed for protein synthesis.

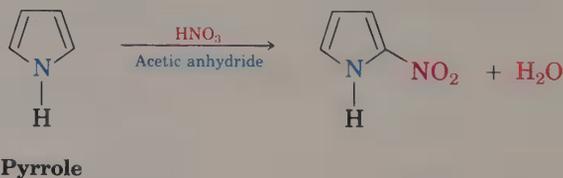
Small DNA segments can be synthesized in the laboratory, and commercial instruments are available for automating the work. Sequencing of DNA can be carried out rapidly and efficiently by the **Maxam-Gilbert method**.

## 29.19 Summary of Reactions

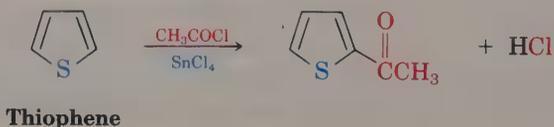
1. Electrophilic aromatic substitution (Section 29.3)
  - a. Bromination



- b. Nitration



- c. Friedel-Crafts acylation





- (d) Transcription (e) Translation (f) Replication (of DNA)  
 (g) Codon (h) Anticodon

- 29.33** What amino acids do these ribonucleotide triplets code for?  
 (a) AAU (b) GAG (c) UCC (d) CAU
- 29.34** From what DNA sequences were each of the mRNA codons in Problem 29.33 transcribed?
- 29.35** What anticodon sequences of tRNA's are coded for by the codons in Problem 29.33?
- 29.36** Draw the complete structure of the ribonucleotide codon U-A-C. For what amino acid does this sequence code?
- 29.37** Draw the complete structure of the deoxyribonucleotide sequence from which the mRNA codon in Problem 29.36 was transcribed.
- 29.38** Give an mRNA sequence that would code for synthesis of met-enkephalin:

H-Tyr-Gly-Gly-Phe-Met-OH

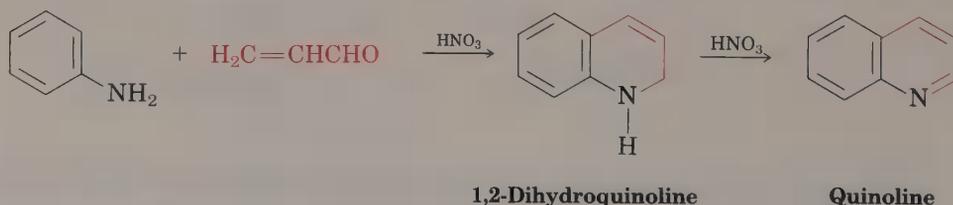
- 29.39** Look up the structure of angiotensin II (Figure 27.4), and give an mRNA sequence that would code for its synthesis.
- 29.40** What amino acid sequence is coded for by this mRNA base sequence?

CUA-GAC-CGU-UCC-AAG-UGA

- 29.41** Isoquinolines are often synthesized by the Bischler-Napieralski cyclization of an *N*-acyl-2-phenylethyl amine with strong acid and  $P_2O_5$ , followed by oxidation of the initially formed dihydroisoquinoline. Suggest a mechanism by which this cyclization occurs.



- 29.42** Quinolines are often prepared by the Skraup<sup>5</sup> synthesis, in which an aniline reacts with an  $\alpha,\beta$ -unsaturated aldehyde and the dihydroquinoline product is oxidized. Suggest a mechanistic pathway for the Skraup reaction.



- 29.43** Show the steps involved in a laboratory synthesis of the DNA fragment with the sequence C-T-A-G.
- 29.44** Review the mechanism shown in Figure 29.15 for the cleavage of deoxyguanosine residues, and propose a mechanism to account for the similar cleavage of deoxyadenosine residues in a DNA chain. Recall that deoxyadenosine is first methylated at N3 prior to hydrolysis.

<sup>5</sup>Hans Zdenko Skraup (1850–1910); b. Austria; professor, Graz, Vienna, Austria.



## Chapter 30

# Orbitals and Organic Chemistry: Pericyclic Reactions

Most organic reactions take place by polar mechanisms, in which an electron-rich nucleophile donates two electrons in forming a bond to an electron-poor electrophile. Many other organic reactions take place by radical mechanisms, in which each of two reactants donates one electron in forming a new bond. Although much remains to be learned about polar and radical reactions, the broad outlines of both classes have been studied for many years and seem relatively well understood.

By contrast, the fundamental principles of *pericyclic reactions*, the third major class of organic reaction mechanisms, have been understood only recently. Numerous individuals have made major contributions, but it was the work of Robert Woodward<sup>1</sup> and Roald Hoffmann<sup>2</sup> in the mid-1960s that made most chemists aware of the principles of pericyclic reactions.

### 30.1 Some Examples of Pericyclic Reactions

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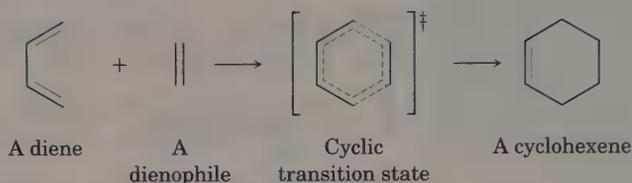
We previously defined a **pericyclic reaction** as one that occurs by a concerted process through a cyclic transition state. The word **concerted** means that all bonding changes occur at the same time and in a single step; no intermediates are involved. Rather than try to expand this definition now, let's look at some examples of the three major classes of pericyclic reactions.

<sup>1</sup>Robert Burns Woodward (1917–1979); b. Boston, Mass.; Ph.D. Massachusetts Institute of Technology (1937); professor, Harvard University (1941–1979); Nobel Prize (1965).

<sup>2</sup>Roald Hoffmann (1937– ); b. Zloczow, Poland; Ph.D. Harvard University (1962); professor, Cornell University; Nobel Prize (1981).

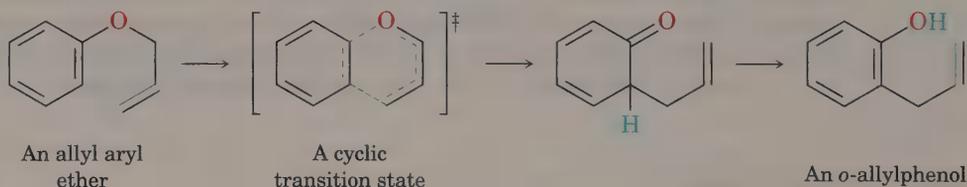
The Diels–Alder reaction (Section 14.7) is an intermolecular pericyclic reaction between a diene and a dienophile to yield a cyclohexene. Three bonds are formed (two sigma bonds and one pi bond) and three bonds are broken (two diene pi bonds and one alkene pi bond) at the same time via a cyclic transition state. This is an example of a **cycloaddition reaction**, so named because the two reagents add together to yield a cyclic product. The product of a cycloaddition always has fewer pi bonds than the starting materials have.

Diels–Alder reaction;  
a cycloaddition



The Claisen rearrangement (Section 26.8) is an intramolecular pericyclic reaction of an allyl aryl ether. Three new bonds are formed (one sigma bond and two pi bonds) and three bonds are broken (one sigma bond and two pi bonds) at the same time via a cyclic transition state. This is an example of a **sigmatropic rearrangement**, so named because the starting molecule rearranges by changing a sigma bond (Greek *tropos*, meaning “change” or “turn”). One sigma bond in the starting material breaks, and a new sigma bond forms in the product.

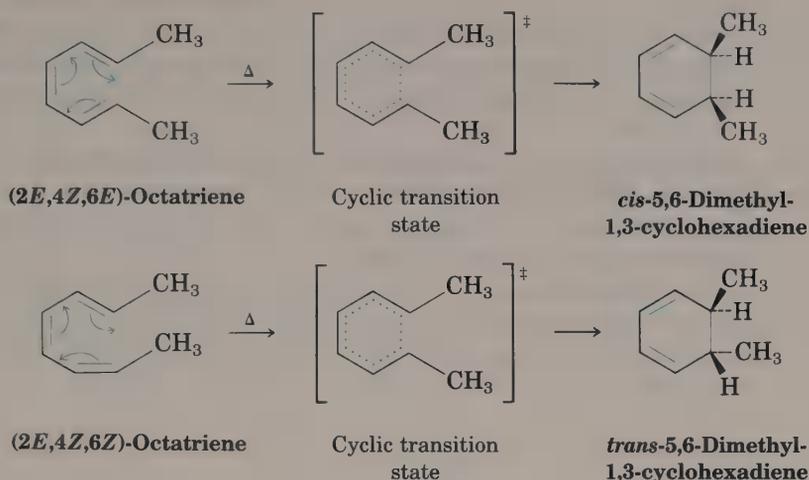
Claisen rearrangement;  
a sigmatropic rearrangement



A third kind of pericyclic process, called an **electrocyclic reaction**, occurs when polyenes are heated or irradiated with ultraviolet light. For example, heating (2*E*,4*Z*,6*E*)-octatriene leads exclusively to formation of *cis*-5,6-dimethyl-1,3-cyclohexadiene. Heating the (2*E*,4*Z*,6*Z*)-octatriene isomer, however, leads exclusively to formation of *trans*-5,6-dimethyl-1,3-cyclohexadiene. These electrocyclic reactions are so named because they involve a cyclic reorganization of electrons. The product of an electrocyclic reaction always has one more sigma bond and one less pi bond than the starting material has.

Both of the above triene cyclizations, and many other pericyclic reactions as well, are *stereospecific*; that is, they yield only a single product stereoisomer rather than a mixture. Why should this be?

## Electrocyclic reactions



The answer to this question is fundamental to the nature of pericyclic reactions and has to do with symmetry properties of the reactant and product orbitals. To understand how orbital symmetry affects reactivity, we need to look more deeply into the nature of molecular orbitals.

## 30.2 Atomic and Molecular Orbitals

Let's begin by briefly reviewing some of the concepts introduced in Chapter 1. We've said that an orbital is the region of space where a given electron is most likely to be found. In mathematical terms, orbitals are derived from quantum mechanical calculations involving electron wave functions. Although the exact nature of these calculations isn't important at the moment, the different lobes of an orbital turn out to have algebraic signs, + and -, when the wave functions are solved. For example, the two equivalent lobes of a  $p$  atomic orbital have plus and minus signs, and the two nonequivalent lobes of an  $sp^3$  hybrid orbital also have plus and minus signs (Figure 30.1).

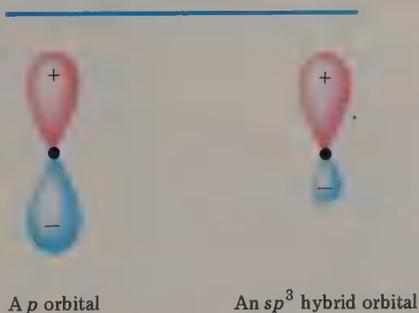
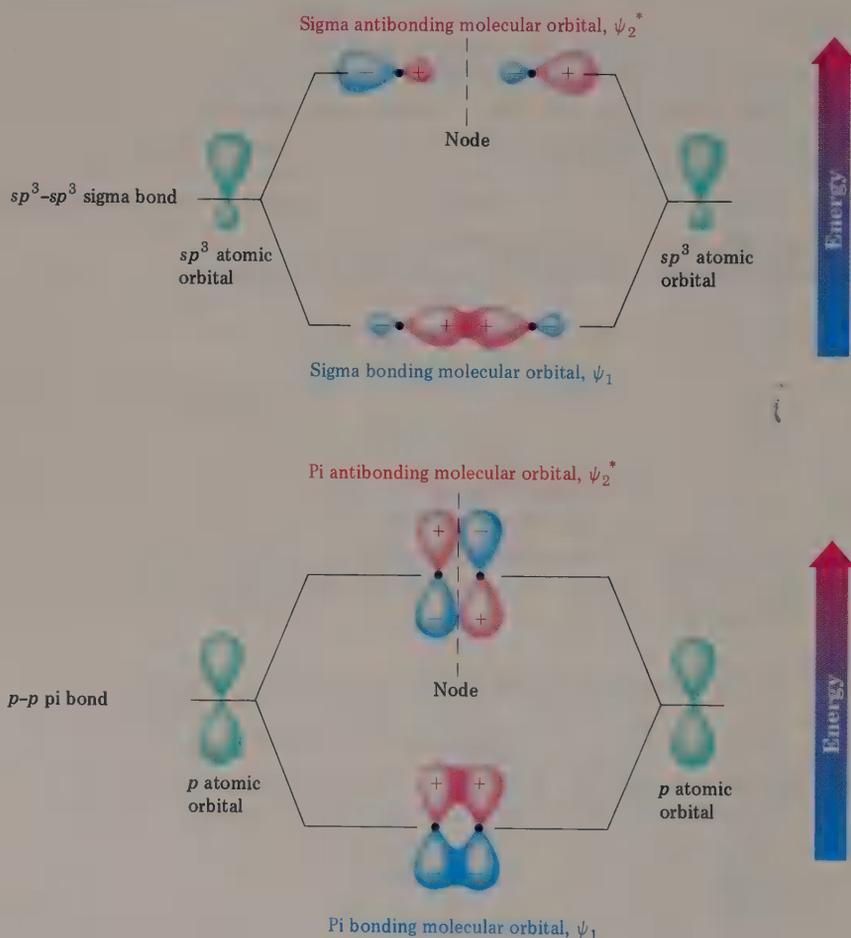


Figure 30.1 Signs of lobes in  $p$  and  $sp^3$  orbitals.

When a bond is formed between two atoms by overlap of atomic orbitals, the two electrons in the bond no longer occupy atomic orbitals but instead occupy molecular orbitals (MO's). Overlap of two atomic orbitals leads to two molecular orbitals, one of which is lower in energy and one of which is higher in energy than either of the original two atomic orbitals.

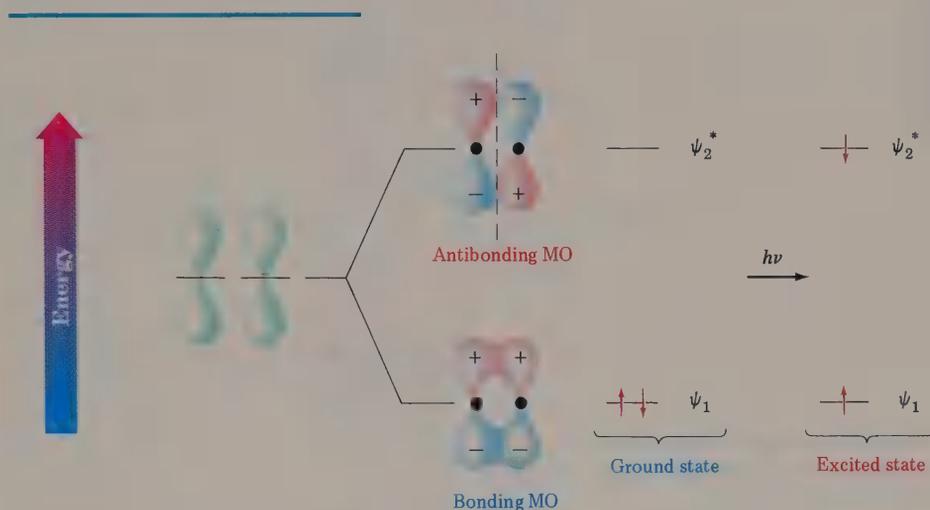
As shown in Figure 30.2, the lower-energy molecular orbital, denoted  $\psi_1$  (Greek psi), is a **bonding molecular orbital** because it results from additive overlap of two lobes with the same algebraic sign. The higher-energy molecular orbital, denoted  $\psi_2^*$ , is an **antibonding molecular orbital** because it results from subtractive overlap of two lobes with the opposite algebraic sign. Note that the antibonding orbitals have nodes between nuclei—planes of zero electron density between lobes of opposite sign. Higher-energy orbitals always have more nodes between nuclei than do lower-energy orbitals and thus have fewer favorable bonding interactions.



**Figure 30.2** Sigma and pi molecular orbitals. Bonding molecular orbitals result from overlap of like lobes, whereas antibonding molecular orbitals result from overlap of unlike lobes.

Having now constructed molecular orbitals, we can arrive at a complete description of sigma and pi bonds by assigning electrons to the orbitals. The method used is exactly the same as that used in assigning electrons to atomic orbitals in describing the electronic configuration of atoms (the *aufbau* principle, Section 1.3). Each molecular orbital can hold two electrons of opposite spin.

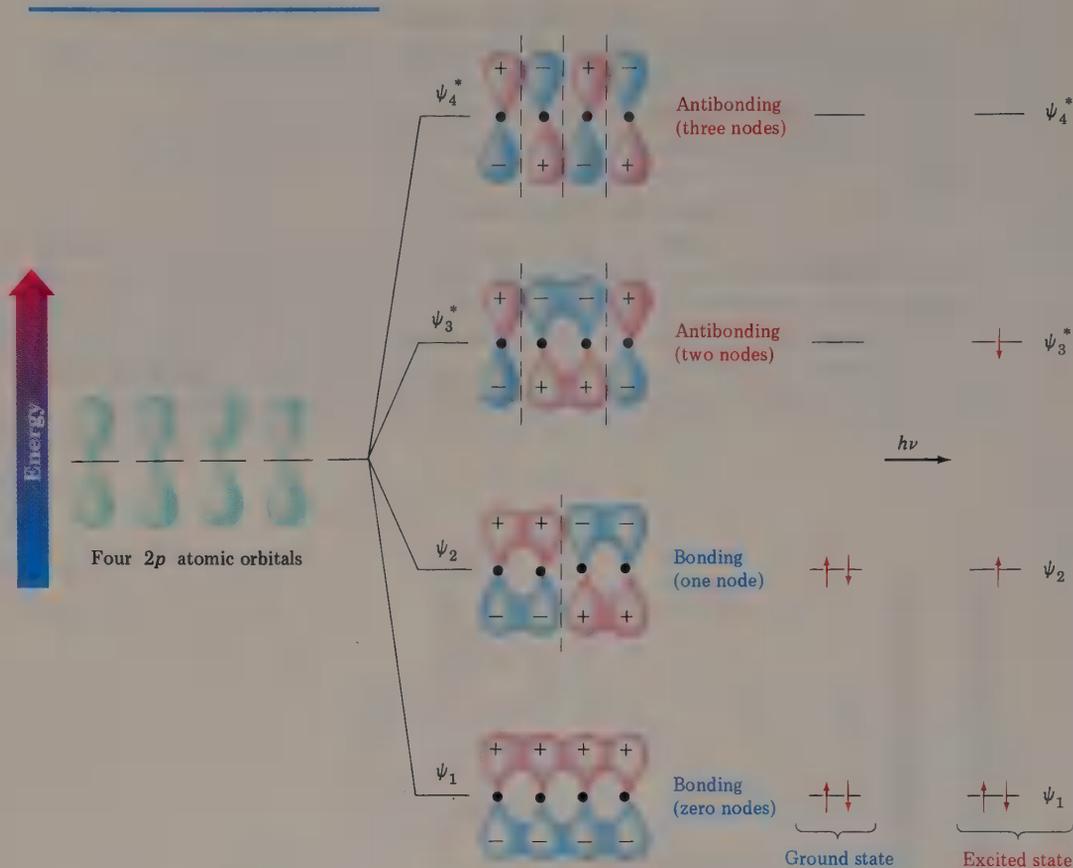
The assignments are made by filling the lowest-energy molecular orbitals first and by filling each molecular orbital with two electrons before going on to the next higher orbital. For example, the ground-state electronic configuration of an alkene pi bond has  $\psi_1$  filled and  $\psi_2^*$  vacant. Irradiation with ultraviolet light, however, excites an electron from  $\psi_1$  to  $\psi_2^*$ , leaving each orbital with one electron (Section 14.11). The electronic configurations of both ground and excited states of ethylene are shown in Figure 30.3.



**Figure 30.3** Ground-state and excited-state electronic configurations of the ethylene pi bond.

### 30.3 Molecular Orbitals of Conjugated PI Systems

The molecular orbital description of a conjugated pi system is more complex than that of a simple alkene because the pi electrons are delocalized over more than two atoms. In 1,3-butadiene, for example, four  $2p$  atomic orbitals combine into four diene molecular orbitals spanning the entire pi system. Two of these molecular orbitals are bonding (lower energy,  $\psi_1$  and  $\psi_2$ ), and two are antibonding (higher energy,  $\psi_3^*$  and  $\psi_4^*$ ), as shown in Figure 30.4. The two bonding orbitals in ground-state butadiene are occupied by four electrons, while the two antibonding orbitals are unoccupied. Note once again that the higher the energy of a molecular orbital, the more nodes it has between nuclei.



**Figure 30.4** Pi molecular orbitals of 1,3-butadiene. In the ground state, the two bonding orbitals are occupied and the two antibonding orbitals are vacant.

A similar sort of molecular orbital description can be derived for a conjugated triene or for *any* conjugated pi electron system. The six molecular orbitals of 1,3,5-hexatriene are shown in Figure 30.5. Only the three bonding orbitals,  $\psi_1$ ,  $\psi_2$ , and  $\psi_3$ , are occupied in the ground state, whereas  $\psi_3$  and  $\psi_4^*$  have one electron each in the excited state.

## 30.4 Molecular Orbitals and Pericyclic Reactions

What do molecular orbitals and the signs of their lobes have to do with pericyclic reactions? The answer is *everything*. According to a series of rules formulated by Woodward and Hoffmann, a pericyclic reaction can take place only if the symmetry of all reactant molecular orbitals is the same as the symmetry of the product molecular orbitals. In other words, the lobes of reactant molecular orbitals must be of the correct algebraic sign for bonding overlap to occur in the transition state leading to the product.

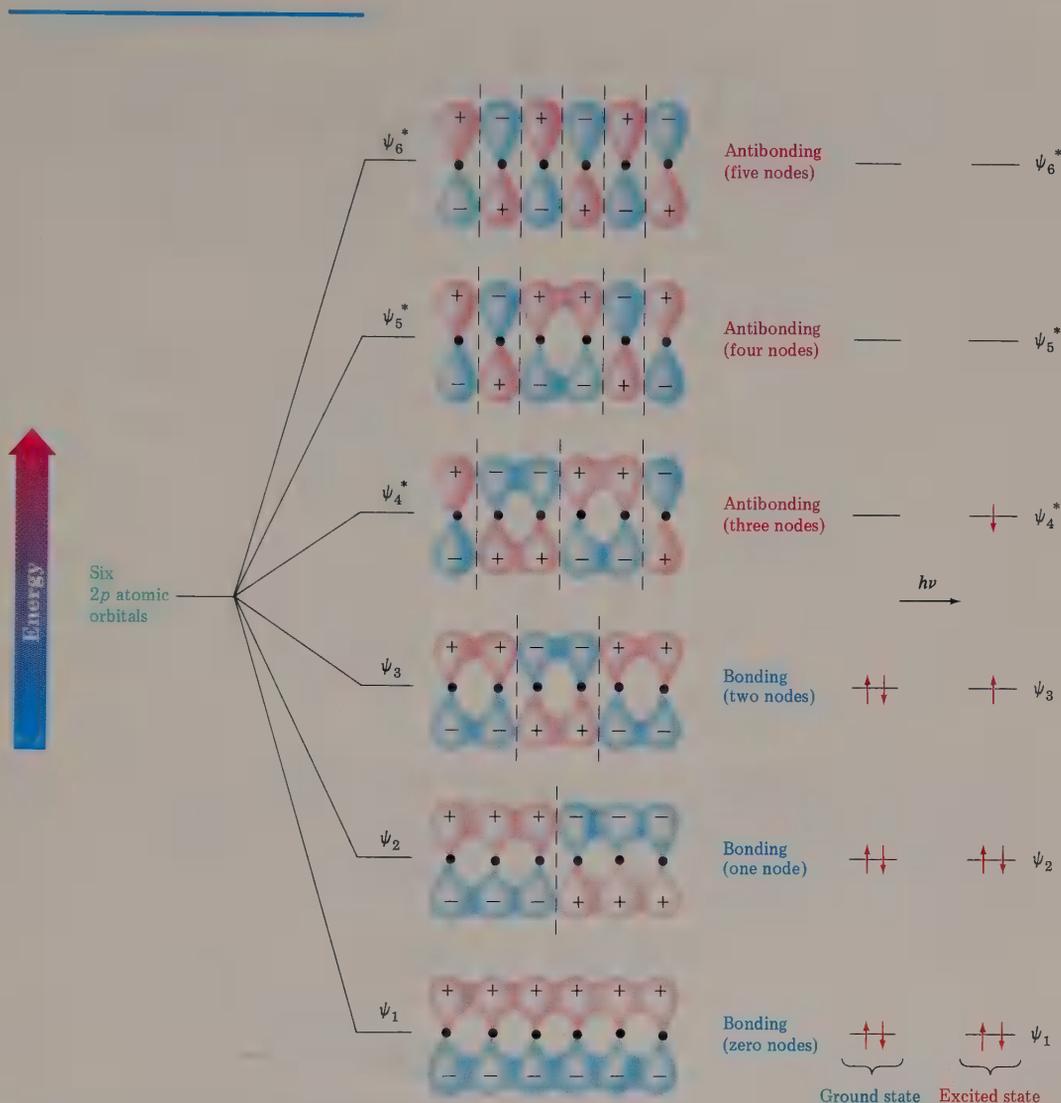


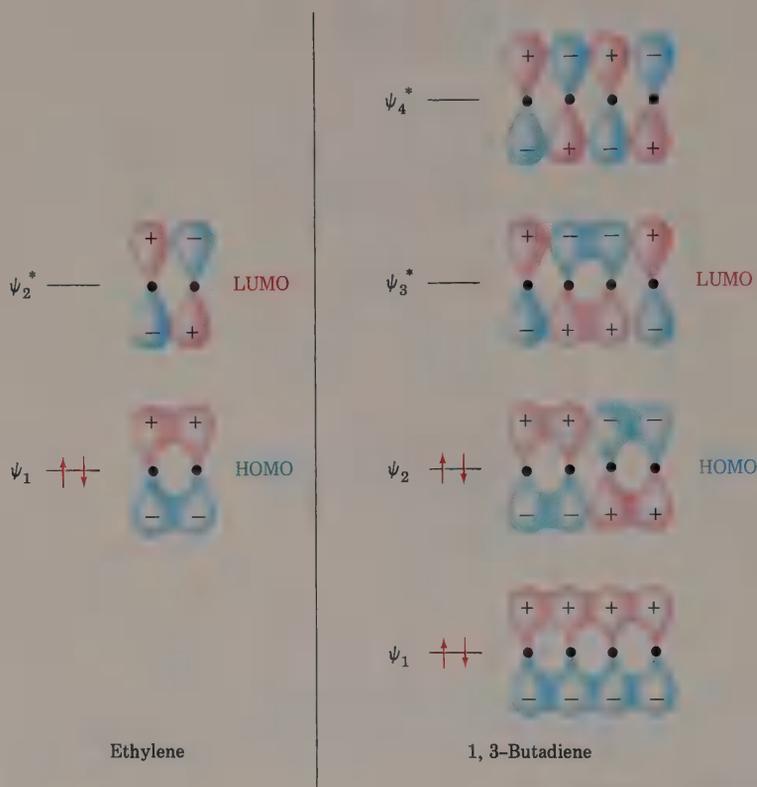
Figure 30.5 Pi molecular orbitals of 1,3,5-hexatriene.

If the orbital symmetries of both reactant and product match up, or “correlate,” the reaction is said to be **symmetry allowed**. If the orbital symmetries of reactant and product don’t correlate, the reaction is **symmetry disallowed**. Symmetry-allowed reactions often occur under relatively mild conditions, but symmetry-disallowed reactions can’t occur by concerted paths. They take place either by nonconcerted pathways under forcing, high-energy conditions, or not at all.

The Woodward–Hoffmann rules for pericyclic reactions require an analysis of *all* reactant and product molecular orbitals, but Kenichi Fukui<sup>3</sup> at

<sup>3</sup>Kenichi Fukui (1918– ); b. Nara Prefecture, Japan; Ph.D. Kyoto University; professor, Kyoto University; Nobel Prize (1981).

Kyoto University in Japan introduced a simplified version. According to Fukui, we need consider only *two* molecular orbitals called the **frontier orbitals**. These frontier orbitals are the **highest occupied molecular orbital (HOMO)** and the **lowest unoccupied molecular orbital (LUMO)**. In ground-state ethylene, for example,  $\psi_1$  is the HOMO since it has two electrons and  $\psi_2^*$  is the LUMO since it is vacant. In ground-state 1,3-butadiene,  $\psi_2$  is the HOMO and  $\psi_3^*$  is the LUMO (Figure 30.6).



**Figure 30.6** The HOMO and LUMO of ground-state ethylene and 1,3-butadiene.

The best way to understand how orbital symmetry affects pericyclic reactions is to look at some examples. Let's look first at some electrocyclic reactions of polyenes and then go on to cycloaddition reactions and sigma-tropic rearrangements.

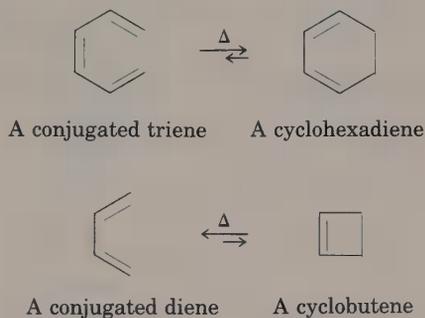
**PROBLEM** .....

- 30.1** Refer to Figure 30.5 to find the molecular orbitals of a conjugated triene, and tell which molecular orbital is the HOMO and which is the LUMO for both ground and excited states.

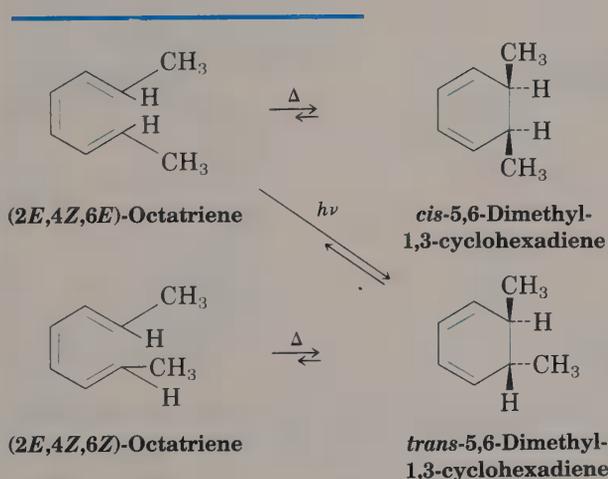
## 30.5 Electrocyclic Reactions

**Electrocyclic reactions** are pericyclic processes that involve the cyclization of conjugated polyenes. One pi bond is broken, the other pi bonds change position, a new sigma bond is formed, and a cyclic compound results. For example, conjugated trienes can be converted into cyclohexadienes, and conjugated dienes can be converted into cyclobutenes.

Both of these reactions are reversible, and the position of the equilibrium depends on the specific case. In general, the triene–cyclohexadiene equilibrium favors the ring-closed product, whereas the diene–cyclobutene equilibrium favors the ring-opened product.

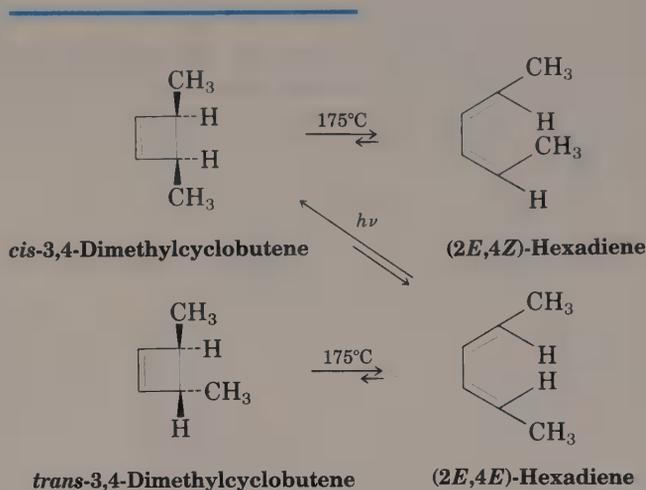


The most important feature of electrocyclic reactions is their stereochemistry. For example, *(2E,4Z,6E)*-octatriene yields only *cis*-5,6-dimethyl-1,3-cyclohexadiene when heated, and *(2E,4Z,6Z)*-octatriene yields only *trans*-5,6-dimethyl-1,3-cyclohexadiene. Remarkably, however, the stereochemical results change completely when the reactions are carried out under *photochemical*, rather than thermal, conditions. Thus, irradiation of *(2E,4Z,6E)*-octatriene with ultraviolet light yields *trans*-5,6-dimethyl-1,3-cyclohexadiene (Figure 30.7).



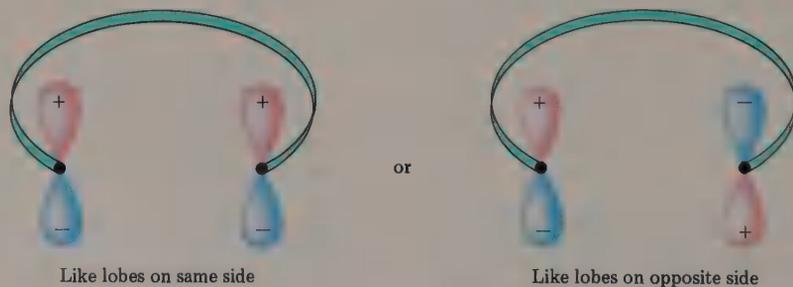
**Figure 30.7** Electrocyclic interconversions of 2,4,6-octatrienes and 5,6-dimethyl-1,3-cyclohexadienes.

A similar result is obtained for the thermal electrocyclic ring opening of the 3,4-dimethylcyclobutenes. The trans isomer yields only (2*E*,4*E*)-hexadiene when heated, and the cis isomer yields only (2*E*,4*Z*)-hexadiene. On UV irradiation, however, the results are again different: Cyclization of the 2*E*,4*E* isomer under photochemical conditions yields cis product (Figure 30.8).

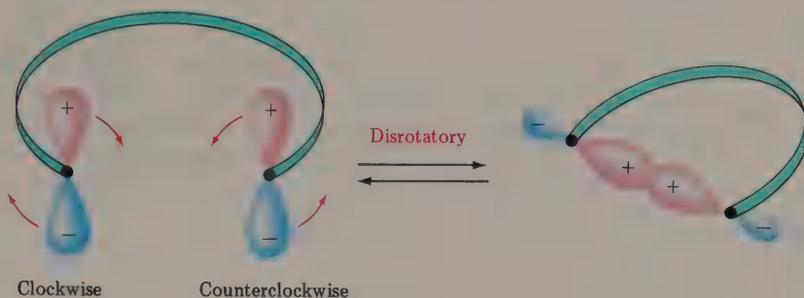


**Figure 30.8** Electrocyclic interconversions of 2,4-hexadienes and 3,4-dimethylcyclobutenes.

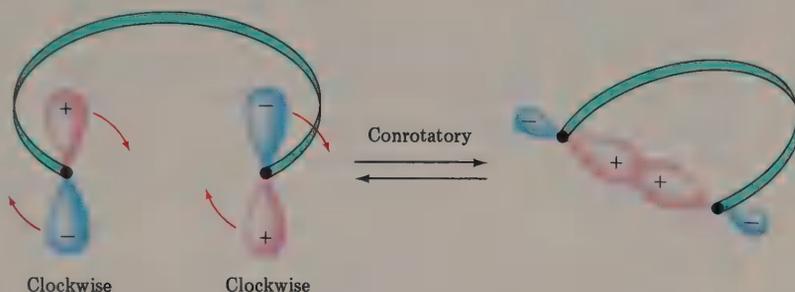
All of these stereospecific electrocyclic reactions can be accounted for by orbital-symmetry arguments. To do so, we need to look only at the symmetries of the two outermost lobes of the polyene. There are two possibilities: The lobes of like sign can be either on the same side or on opposite sides of the molecule.



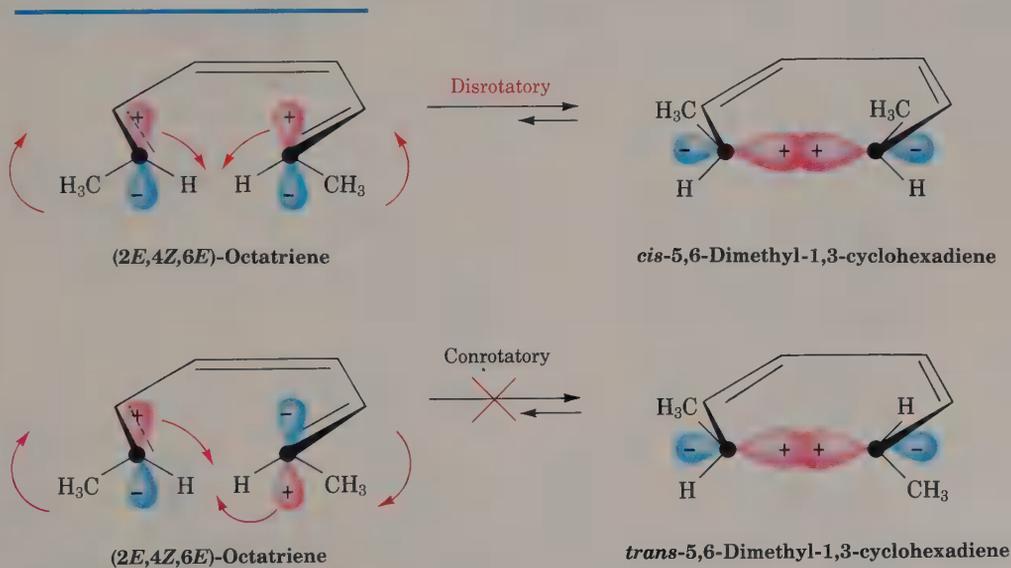
For a bond to form, the outermost pi lobes must rotate so that favorable bonding overlap is achieved—a positive lobe overlapping a positive lobe or a negative lobe overlapping a negative lobe. If two lobes of like sign are on the same side of the molecule, the two orbitals must rotate in *different* directions. One orbital must rotate clockwise, and one must rotate counter-clockwise. This kind of motion is referred to as **disrotatory**:



Conversely, if lobes of like sign are on opposite sides of the molecules, both orbitals must rotate in the *same* direction, either both clockwise or both counterclockwise. This kind of motion is called **conrotatory**:



Now let's see what happens to substituents on the polyene carbon atoms when cyclization occurs. To choose a specific example, let's look at the thermal cyclization of (2*E*,4*Z*,6*E*)-octatriene. If a disrotatory cyclization were to occur, *cis*-5,6-dimethyl-1,3-cyclohexadiene would result. If a conrotatory cyclization were to occur, *trans*-5,6-dimethyl-1,3-cyclohexadiene would result (Figure 30.9).



**Figure 30.9** Stereochemistry of disrotatory and conrotatory cyclization of (2*E*,4*Z*,6*E*)-octatriene.

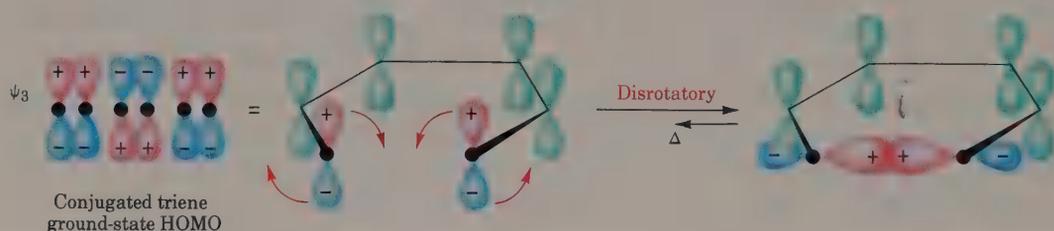
In fact, only the disrotatory cyclization of  $(2E,4Z,6E)$ -octatriene is observed to occur. We therefore conclude that the stereochemistry of an electrocyclic reaction is determined by the mode of ring closure. The mode of ring closure, in turn, is determined by the symmetry of reactant molecular orbitals.

## 30.6 Stereochemistry of Thermal Electrocyclic Reactions

How can we predict which mode of ring closure, conrotatory or disrotatory, will occur in a given case? How can we tell whether the terminal lobes of like sign will be on the same side or on opposite sides of the molecule?

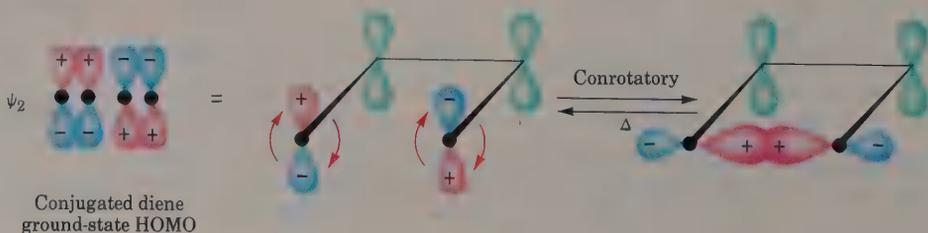
According to frontier orbital theory, *the stereochemistry of an electrocyclic reaction is determined by the symmetry of the polyene's HOMO*. The electrons in the HOMO are the highest-energy, most loosely held electrons, and are therefore most easily moved during reaction. For thermal ring openings and closings, the ground-state electronic configuration is used to identify the HOMO; for photochemical ring openings and closings, the excited-state electronic configuration is used.

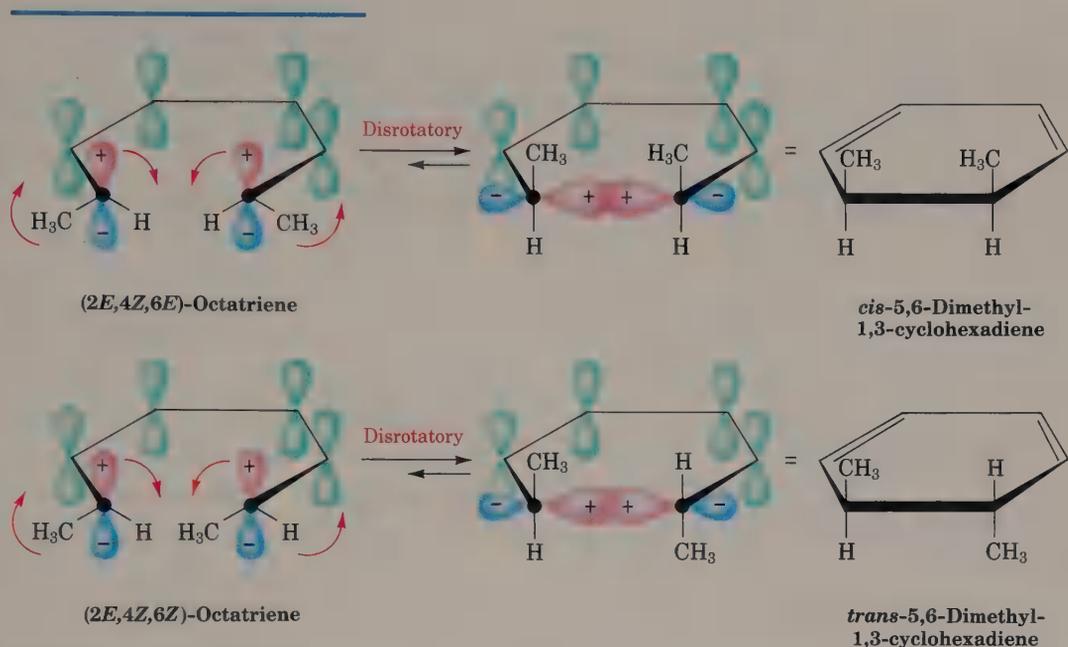
Let's look again at the thermal ring closure of conjugated trienes. According to Figure 30.5, the HOMO of a conjugated triene in its ground state has a symmetry that predicts a disrotatory ring closure:



This disrotatory cyclization is exactly what's observed in the thermal cyclization of 2,4,6-octatriene. The  $2E,4Z,6E$  isomer yields *cis* product; the  $2E,4Z,6Z$  isomer yields *trans* product (Figure 30.10).

In a similar manner, the ground-state HOMO of a conjugated diene (Figure 30.4) has a symmetry that predicts conrotatory ring closure:

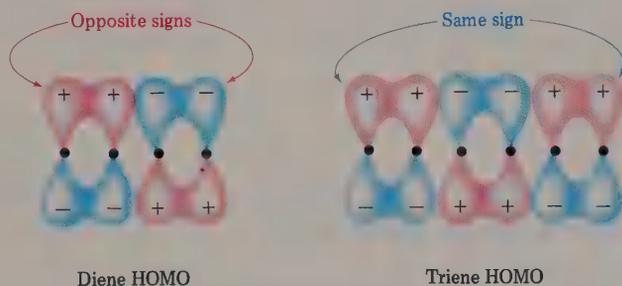




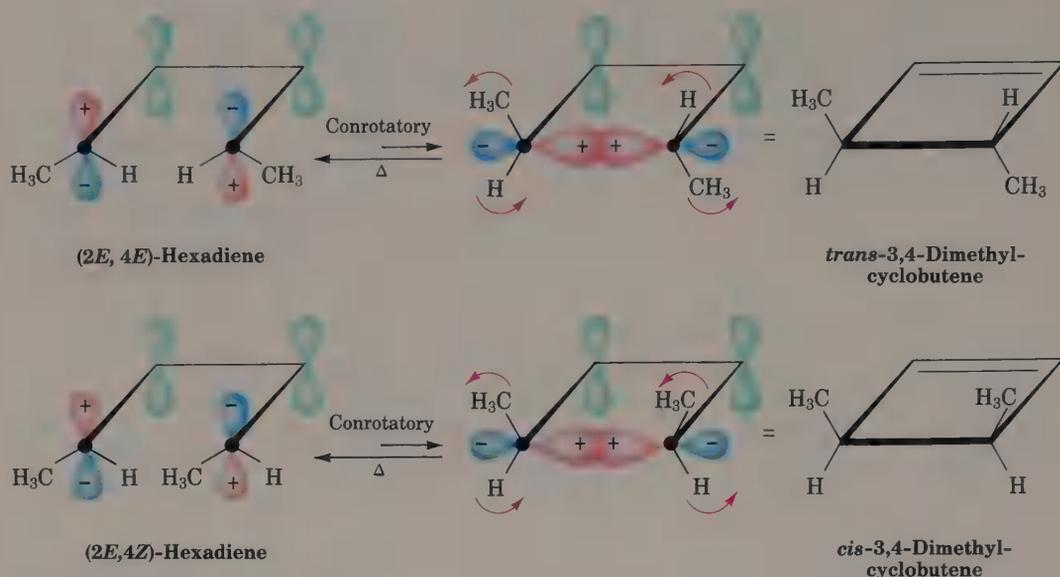
**Figure 30.10** Thermal disrotatory ring closure of 2,4,6-octatrienes.

In practice, of course, the conjugated diene reaction can only be observed in the reverse direction (cyclobutene  $\rightarrow$  butadiene) because of the position of the equilibrium. We therefore predict that the 3,4-dimethylcyclobutene ring will *open* in a conrotatory fashion, which is exactly what is observed. *cis*-3,4-Dimethylcyclobutene yields (2*E*,4*Z*)-hexadiene, and *trans*-3,4-dimethylcyclobutene yields (2*E*,4*E*)-hexadiene by conrotatory opening (Figure 30.11 on page 1162).

Note that a conjugated diene and a conjugated triene react in opposite stereochemical senses. The diene opens and closes by a conrotatory path, whereas the triene opens and closes by a disrotatory path. This difference, of course, is due to the different symmetries of the diene and triene HOMO's:



There is an alternating relationship between the number of electron pairs (double bonds) undergoing bond reorganization and the mode of ring opening or closure. Polyenes with an even number of electron pairs undergo thermal electrocyclic reactions in a conrotatory sense, whereas polyenes



**Figure 30.11** Conrotatory ring opening of the *cis*- and *trans*-dimethylcyclobutenes.

with an odd number of electron pairs undergo the same reactions in a disrotatory sense.

**PROBLEM** .....

- 30.2** Draw the products you would expect to obtain from conrotatory and disrotatory cyclizations of (2*Z*,4*Z*,6*Z*)-octatriene. Which of the two paths would you expect the thermal reaction to follow?

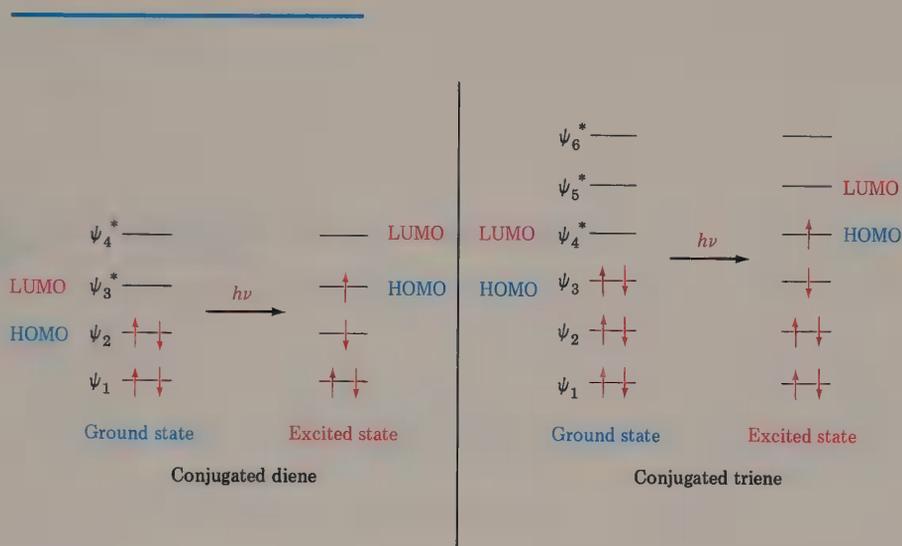
**PROBLEM** .....

- 30.3** In principle, *trans*-3,4-dimethylcyclobutene can open by two conrotatory paths to give either (2*E*,4*E*)-hexadiene or (2*Z*,4*Z*)-hexadiene. Explain why both products are symmetry allowed, and then account for the fact that only the 2*E*,4*E* isomer is obtained in practice.
- .....

## 30.7 Photochemical Electrocyclic Reactions

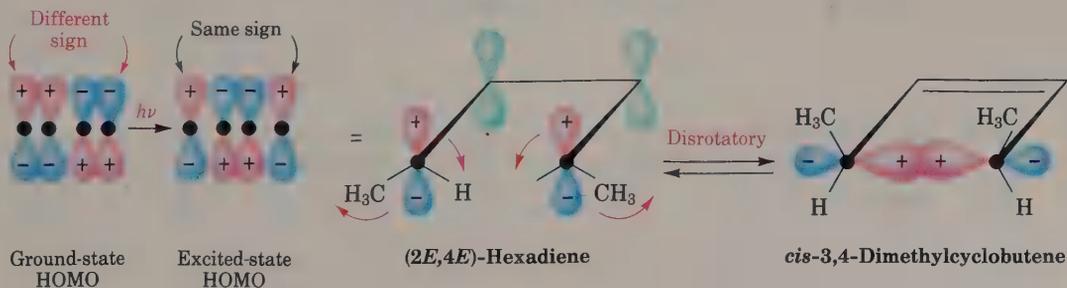
We noted previously that photochemical electrocyclic reactions take a different stereochemical course than their thermal counterparts do. We can now explain this difference. Ultraviolet irradiation of a polyene causes an

excitation of one electron from the ground-state HOMO to the ground-state LUMO. For example, irradiation of a conjugated diene excites an electron from  $\psi_2$  to  $\psi_3^*$ , and irradiation of a conjugated triene excites an electron from  $\psi_3$  to  $\psi_4^*$  (Figure 30.12).

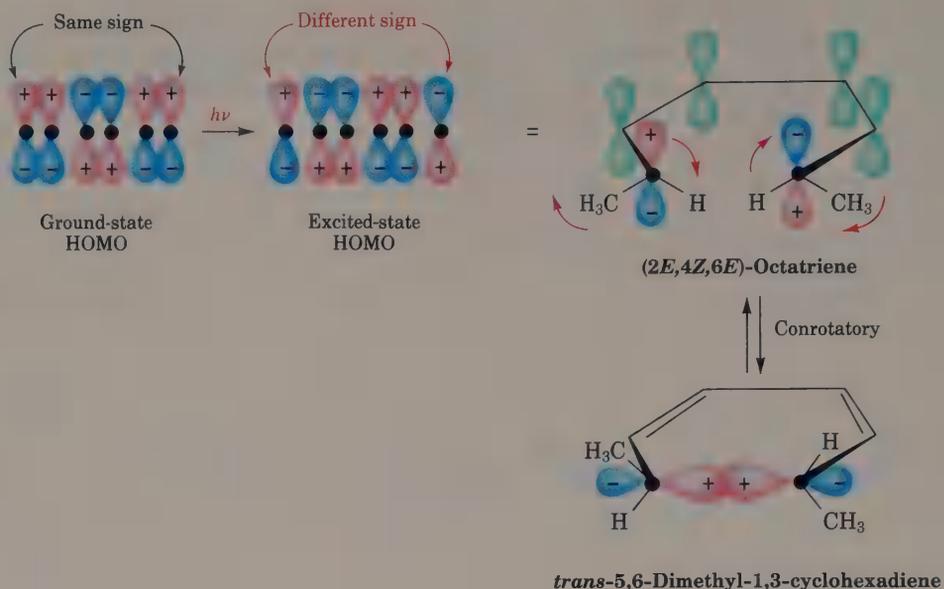


**Figure 30.12** Ground-state and excited-state electronic configurations of conjugated dienes and trienes.

Electronic excitation changes the symmetries of HOMO and LUMO and hence changes the reaction stereochemistry. Thus,  $(2E,4E)$ -hexadiene undergoes *photochemical* cyclization by a disrotatory path (recall that the thermal reaction is conrotatory):



In the same way,  $(2E,4Z,6E)$ -octatriene undergoes photochemical cyclization by a conrotatory path (the thermal reaction is disrotatory):



Thermal and photochemical electrocyclic reactions *always* take place with opposite stereochemistry, as the rules governing these processes indicate (Table 30.1). Learning these simple rules makes it possible to predict the stereochemistry of a large number of organic reactions.

**TABLE 30.1** Stereochemical Rules for Electrocyclic Reactions

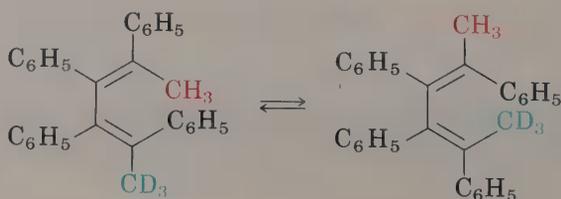
<i>Electron pairs (double bonds)</i>	<i>Thermal reaction</i>	<i>Photochemical reaction</i>
Even number	Conrotatory	Disrotatory
Odd number	Disrotatory	Conrotatory

**PROBLEM** .....

- 30.4** What product would you expect to obtain from the photochemical cyclizations of (2*E*,4*Z*,6*E*)-octatriene? Of (2*E*,4*Z*,6*Z*)-octatriene?

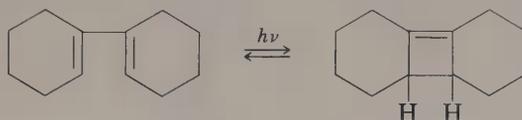
**PROBLEM** .....

- 30.5** The following thermal isomerization has been reported to occur under relatively mild conditions. Identify the pericyclic reactions involved, and show how the rearrangement occurs.



## PROBLEM

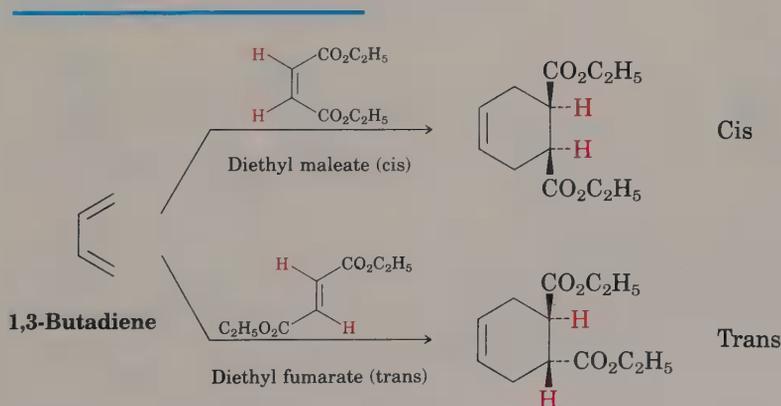
- 30.6 Would you expect the following reaction to proceed in a conrotatory or disrotatory manner? Show the stereochemistry of the cyclobutene product, and explain your answer.



## 30.8 Cycloaddition Reactions

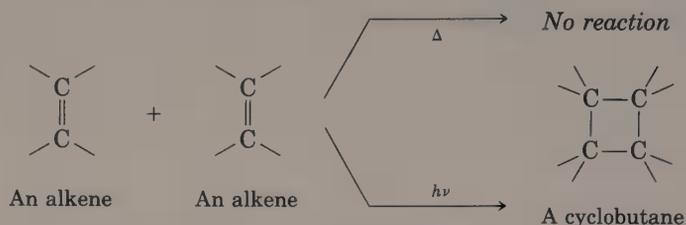
**Cycloaddition reactions** are intermolecular pericyclic processes in which two unsaturated molecules add to each other, yielding a cyclic product. As with electrocyclic reactions, cycloadditions are controlled by the orbital symmetry of the reactants. Symmetry-allowed processes often take place readily, but symmetry-disallowed processes take place with great difficulty, if at all, and then only by nonconcerted pathways. Let's look at two possible reactions to see how they differ.

The Diels–Alder cycloaddition reaction is a [4 + 2]- $\pi$ -electron process that takes place between a diene (4  $\pi$  electrons) and a dienophile (2  $\pi$  electrons) to yield a cyclohexene product. Thousands of examples of Diels–Alder reactions are known. They often take place under mild conditions (room temperature or slightly above), and they are stereospecific with respect to substituents. For example, room-temperature reaction between 1,3-butadiene and diethyl maleate (*cis*) yields exclusively *cis*-disubstituted cyclohexene product; reaction between 1,3-butadiene and diethyl fumarate (*trans*) yields exclusively *trans*-disubstituted product (Figure 30.13).

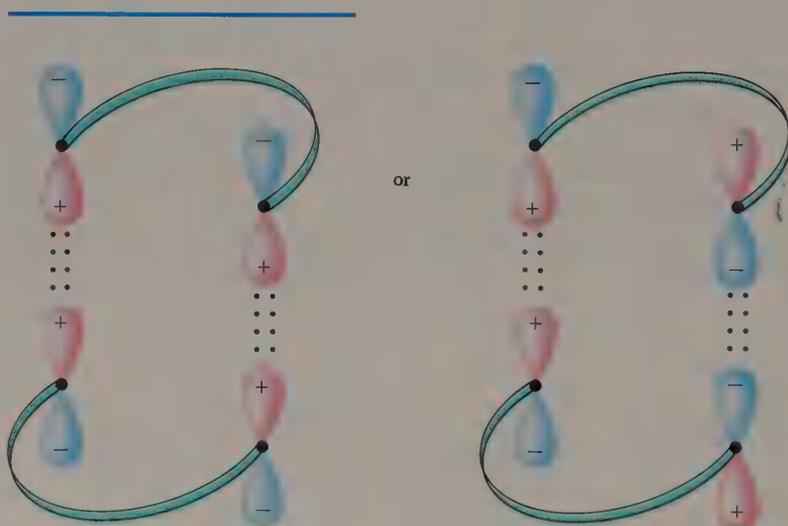


**Figure 30.13** Diels–Alder cycloaddition reactions of diethyl maleate (*cis*) and diethyl fumarate (*trans*). The reactions are stereospecific.

In contrast to the [4 + 2]-pi-electron Diels–Alder reaction, thermal cycloaddition between two alkenes (2 pi + 2 pi) does not occur. Photochemical [2 + 2] cycloadditions often take place readily, however, to yield cyclobutane products. How can we use orbital-symmetry arguments to explain these results?

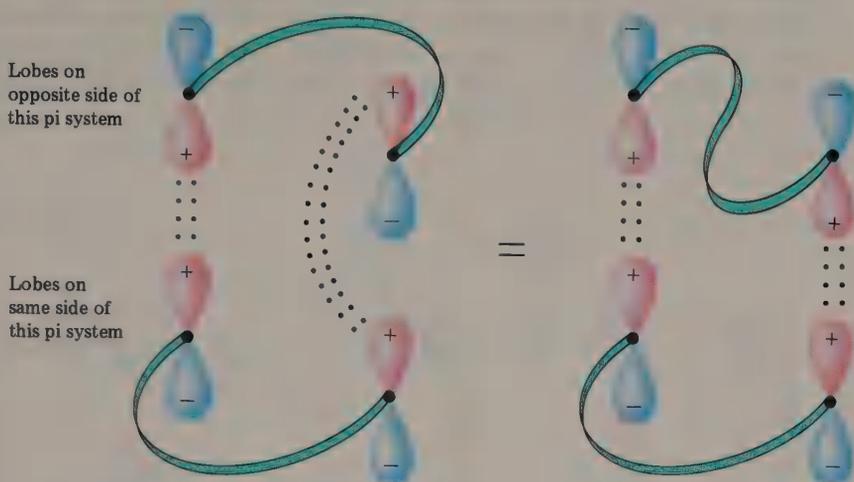


For a successful cycloaddition to take place, the terminal pi lobes of the two unsaturated reactants must have the correct symmetry for bonding overlap to occur. This can happen in either of two ways, designated *suprafacial* and *antarafacial*. **Suprafacial** cycloadditions take place when the orbital symmetries of the reactants are such that overlap occurs between lobes on the *same* face of one component and lobes on the *same* face of the other component (Figure 30.14).



**Figure 30.14** Suprafacial cycloaddition occurs when there is bonding overlap between lobes on the same face of one reactant and lobes on the same face of a second reactant.

**Antarafacial** cycloadditions take place when the orbital symmetries of the reactants are such that overlap occurs between lobes on the *same* face of one component and lobes on *opposite* faces of the other component (Figure 30.15).



**Figure 30.15** Antarafacial cycloaddition occurs when there is bonding overlap between lobes on the same face of one reactant and lobes on opposite faces of the second reactant.

Note that both suprafacial and antarafacial cycloadditions are allowed on orbital-symmetry grounds. Geometric constraints often make antarafacial reactions difficult, however, since there must be twisting of the  $p$  orbital system. Thus, only suprafacial cycloadditions are possible for small pi systems.

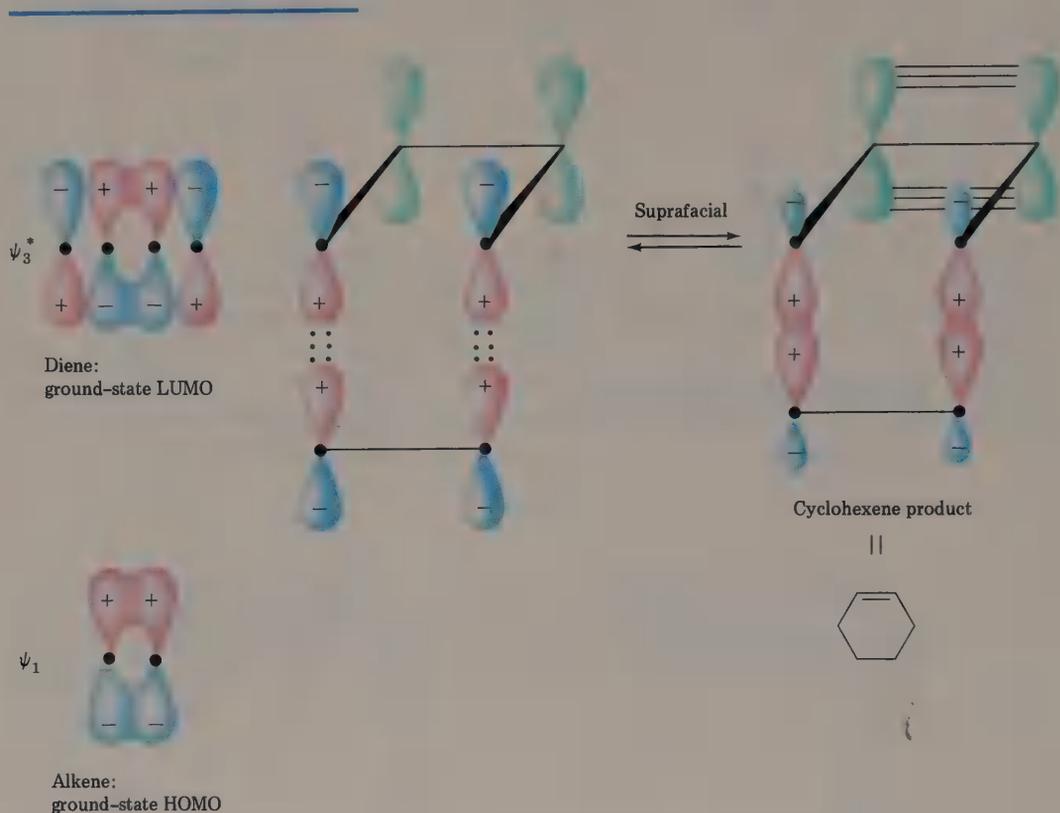
## 30.9 Stereochemistry of Cycloadditions

How can we predict whether a given reaction will occur with suprafacial or with antarafacial geometry? According to frontier orbital theory, cycloaddition reactions take place when the HOMO of one reaction partner overlaps the LUMO of the other reaction partner in a bonding manner.

A good intuitive explanation of this rule is to imagine that one partner reacts by donating two electrons to the second partner. As with electrocyclic reactions, it's the electrons in the HOMO of the first partner that are least tightly held and most likely to be donated. Since only two electrons can be in any one orbital, these electrons must go into a vacant orbital of the second partner, and it is the LUMO that's lowest in energy and most likely to accept the electrons. Let's see how this rule applies to specific cases.

For the  $[4 + 2]$ -pi-electron cycloaddition (Diels–Alder reaction), let's arbitrarily select the diene LUMO and the alkene HOMO. (We could equally well use the diene HOMO and the alkene LUMO.) The symmetries of the two orbitals are such that bonding overlap of the terminal lobes can occur

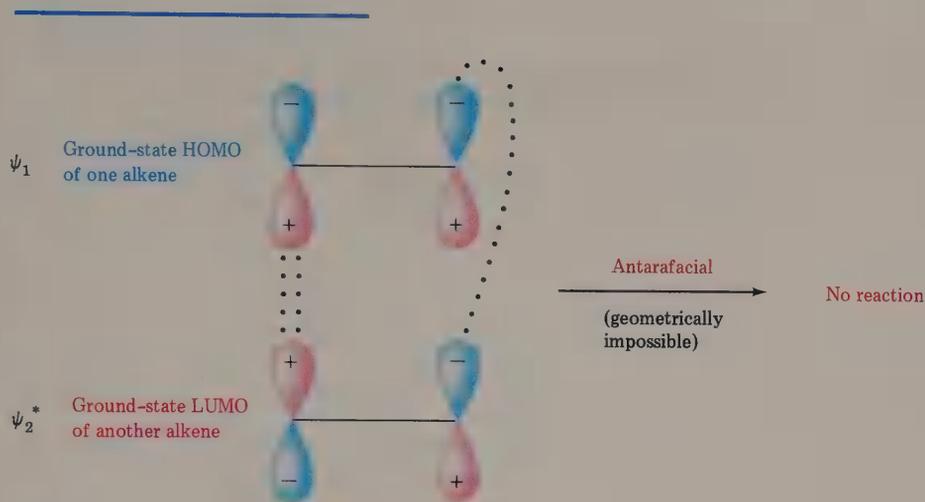
with suprafacial geometry (Figure 30.16). The Diels–Alder reaction therefore takes place readily under thermal conditions. Note that, as with electrocyclic reactions, we need be concerned only with the symmetries of the *terminal* lobes. For purposes of prediction, it doesn't matter whether or not the interior lobes have bonding or antibonding overlap in the product.



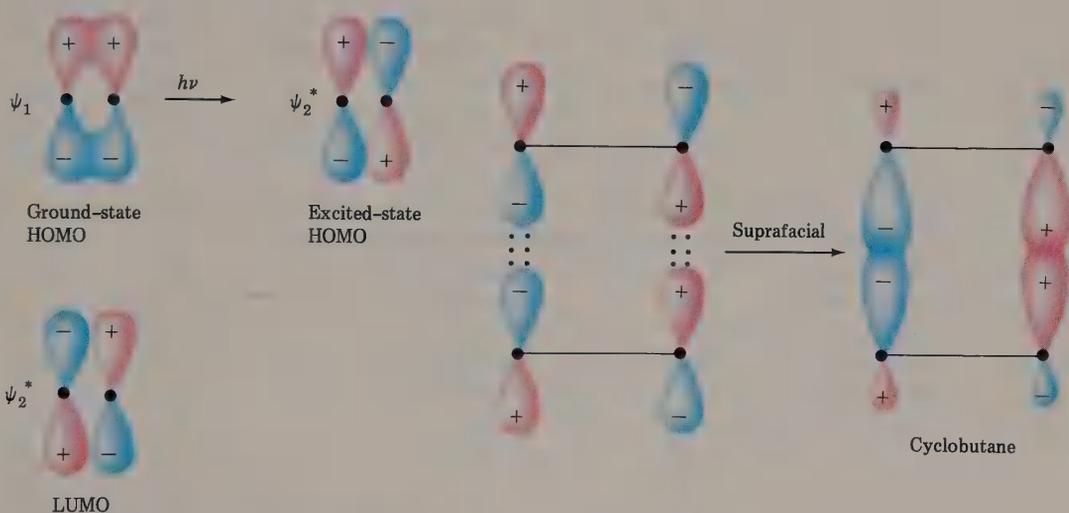
**Figure 30.16** Correlation of diene LUMO and alkene HOMO in a suprafacial [4 + 2] cycloaddition reaction (Diels–Alder reaction).

In contrast to the [4 + 2] Diels–Alder reaction, the [2 + 2] cycloaddition of two alkenes to yield a cyclobutane does not occur thermally but can only be observed photochemically. The explanation follows from orbital-symmetry arguments. Looking at the HOMO of one alkene and the LUMO of the second alkene, it's apparent that a thermal [2 + 2] cycloaddition must take place by an antarafacial pathway (Figure 30.17). Geometric constraints make the antarafacial transition state difficult, however, and concerted thermal [2 + 2] cycloadditions are therefore rarely observed.

In contrast to the thermal case, photochemical [2 + 2] cycloadditions *are* observed. Irradiation of an alkene with ultraviolet light excites an electron from  $\psi_1$ , the ground-state HOMO, to  $\psi_2^*$ , the excited-state HOMO. Correlation between the excited-state HOMO of one alkene and the LUMO of the second alkene indicates that a photochemical [2 + 2] cycloaddition reaction can occur by a suprafacial pathway (Figure 30.18).

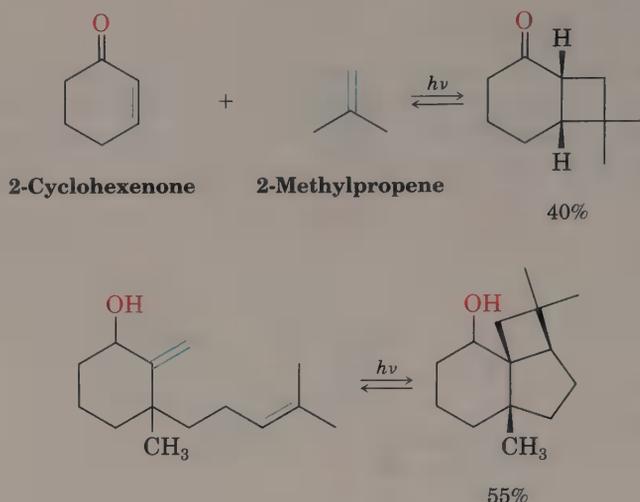


**Figure 30.17** Correlation of HOMO and LUMO in thermal [2 + 2] cycloaddition. The reaction rarely occurs because antarafacial geometry is too strained.



**Figure 30.18** Correlation of HOMO and LUMO in photochemical [2 + 2] cycloaddition reactions. The reaction occurs with suprafacial geometry.

The photochemical [2 + 2] cycloaddition reaction occurs smoothly and represents one of the best methods known for synthesizing cyclobutane rings. The reaction can take place either inter- or intramolecularly, as the following examples show:



Thermal and photochemical cycloaddition reactions always take place by opposite stereochemical pathways. As with electrocyclic reactions, we can categorize cycloadditions according to the total number of electron pairs (double bonds) involved in the rearrangement. Thus, a Diels–Alder [4 + 2] reaction between a diene and a dienophile involves an odd number (three) of electron pairs and takes place by a ground-state suprafacial pathway. A [2 + 2] reaction between two alkenes involves an even number (two) of electron pairs and takes place by a ground-state antarafacial pathway. The general rules governing such processes are given in Table 30.2.

**TABLE 30.2** Stereochemical Rules for Cycloaddition Reactions

<i>Electron pairs (double bonds)</i>	<i>Thermal reaction</i>	<i>Photochemical reaction</i>
Even number	Antarafacial	Suprafacial
Odd number	Suprafacial	Antarafacial

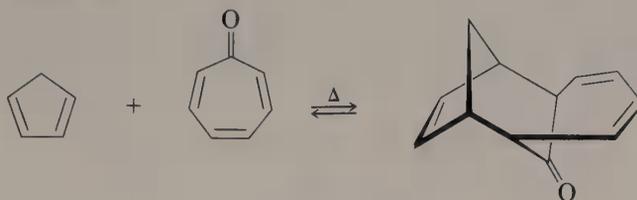
It should be reiterated that both suprafacial and antarafacial cycloaddition pathways are symmetry allowed. Only the geometric constraints inherent in twisting a conjugated pi electron system out of planarity make antarafacial reaction geometry difficult in many cases.

**PROBLEM** .....

- 30.7** What stereochemistry would you expect for the product of the Diels–Alder reaction between (2*E*,4*E*)-hexadiene and ethylene? What stereochemistry would you expect if (2*E*,4*Z*)-hexadiene were used instead?

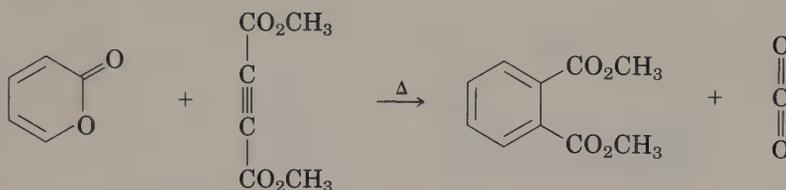
## PROBLEM .....

- 30.8 Cyclopentadiene reacts with cycloheptatrienone to give the product shown. Tell what kind of reaction is involved, and explain the observed result. Is the reaction suprafacial or antarafacial?



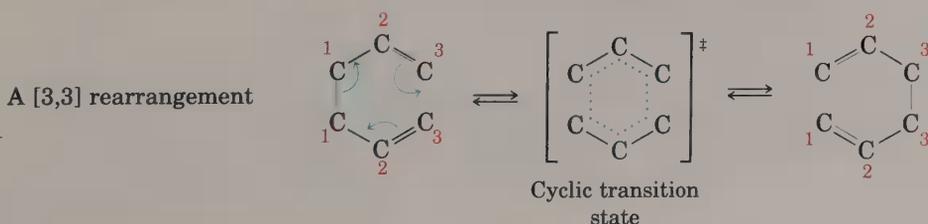
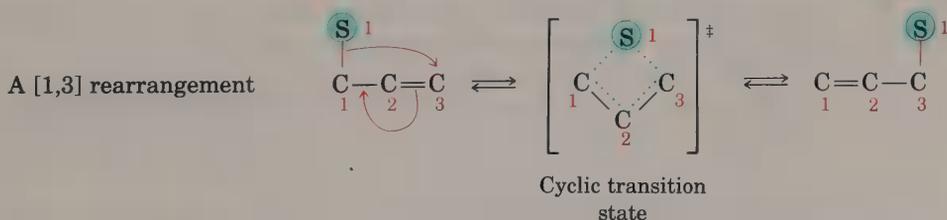
## PROBLEM .....

- 30.9 The following reaction takes place in two steps, one of which is a cycloaddition and the other of which is a *reverse* cycloaddition. Identify the two pericyclic reactions, and show how they occur.



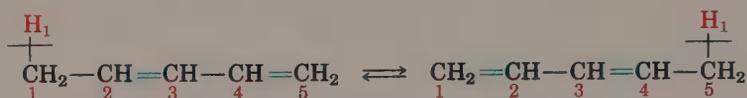
## 30.10 Sigmatropic Rearrangements

**Sigmatropic rearrangements** are pericyclic reactions in which a sigma-bonded substituent group (denoted here by a circled S) migrates across a pi electron system. One sigma bond is broken in the starting material, and a new sigma bond is formed in the product. The sigma-bonded group can be either at the end or in the middle of the pi system, as the following [1,3] and [3,3] rearrangements illustrate:



The designations [1,3] and [3,3] describe the kind of rearrangement that has occurred. The two numbers in brackets refer to the two groups connected by the sigma bond and designate the positions in those groups *to which migration occurs*. For example, in the [1,5] sigmatropic rearrangement of a diene, the two groups connected by the sigma bond are a hydrogen atom and a pentadienyl fragment. Migration occurs to position 1 of the H group (the only possibility) and to position 5 of the pentadienyl group.

A [1,5] sigmatropic rearrangement



In the [3,3] Claisen rearrangement, the two groups connected by the sigma bond are an allyl group and a vinylic ether group. Migration occurs to position 3 of the allyl group and also to position 3 of the vinylic ether.

Claisen rearrangement,  
a [3,3] rearrangement



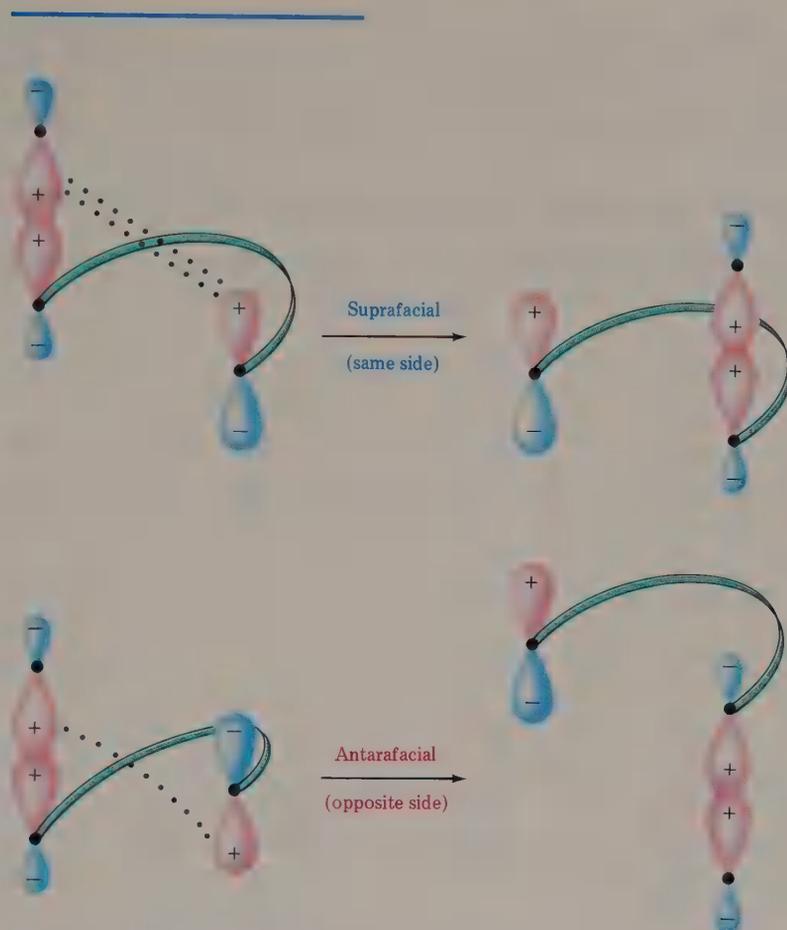
## 30.11 Stereochemistry of Sigmatropic Rearrangements

Sigmatropic rearrangements are more complex than either electrocyclic or cycloaddition reactions but are nonetheless controlled by orbital-symmetry considerations. There are two possible modes of reaction: Migration of a group across the same face of the pi system is called a suprafacial rearrangement, and migration of a group from one face of the pi system to the other face is an antarafacial rearrangement (Figure 30.19).

The rules for sigmatropic rearrangements are identical to those for cycloaddition reactions, as summarized in Table 30.3. Both suprafacial and

**TABLE 30.3** Stereochemical Rules for Sigmatropic Rearrangements

<i>Electron pairs (double bonds)</i>	<i>Thermal reaction</i>	<i>Photochemical reaction</i>
Even number	Antarafacial	Suprafacial
Odd number	Suprafacial	Antarafacial

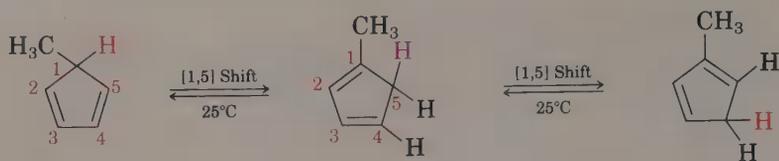


**Figure 30.19** Suprafacial and antarafacial sigmatropic rearrangements.

antarafacial sigmatropic rearrangements are symmetry-allowed processes, but suprafacial rearrangements are often easier for geometric reasons.

## 30.12 Some Examples of Sigmatropic Rearrangements

Since a [1,5] sigmatropic rearrangement involves three electron pairs (two pi bonds and one sigma bond), the orbital-symmetry rules in Table 30.3 predict a suprafacial reaction. In fact, the [1,5] suprafacial shift of a hydrogen atom across two double bonds of a pi system is one of the most commonly observed of all sigmatropic rearrangements. For example, 5-methylcyclopentadiene rapidly rearranges at room temperature to yield a mixture of 1-methyl-, 2-methyl-, and 5-methyl-substituted products.

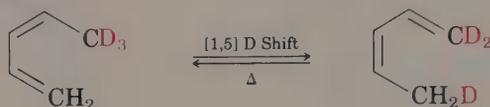


5-Methylcyclopentadiene

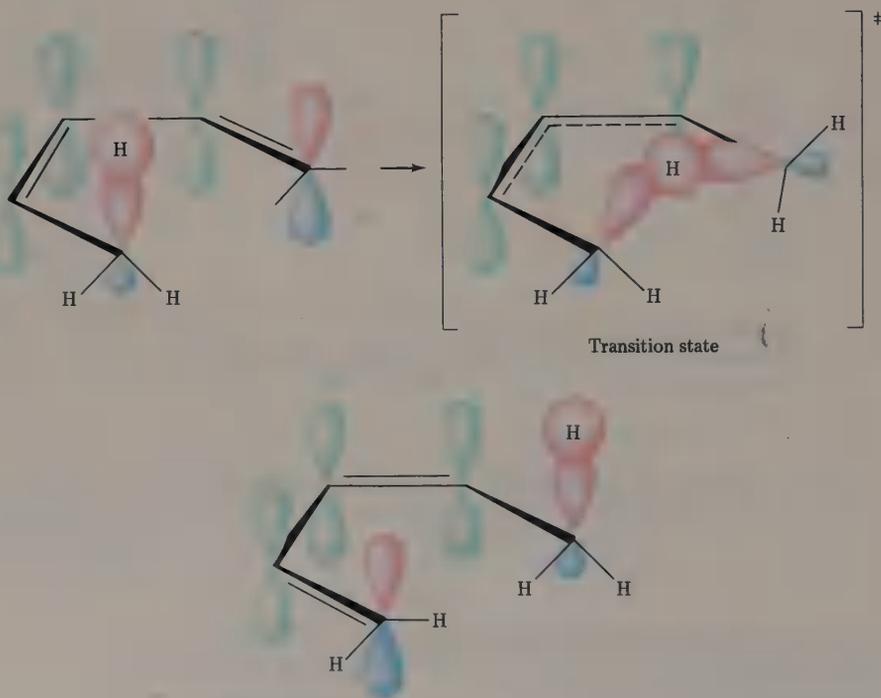
1-Methylcyclopentadiene

2-Methylcyclopentadiene

As another example, heating 5,5,5-trideuterio-(1,3*Z*)-pentadiene causes scrambling of deuterium between positions 1 and 5.



Both of these [1,5] hydrogen shifts occur by a symmetry-allowed suprafacial rearrangement, as illustrated in Figure 30.20.



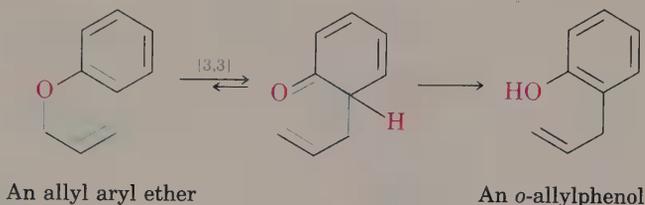
**Figure 30.20** An orbital view of a suprafacial [1,5] hydrogen shift.

In contrast to the preceding two examples of [1,5] sigmatropic hydrogen shifts, thermal [1,3] hydrogen shifts are unknown. Were they to occur, they would have to proceed via an impossibly strained antarafacial reaction pathway.

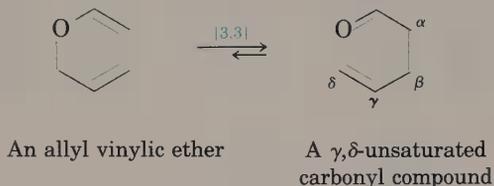
Two other important sigmatropic reactions are the **Cope rearrangement** of a 1,5-hexadiene and the **Claisen rearrangement** of an allyl aryl

ether (Section 26.8). These two, along with the Diels–Alder reaction, are the most useful pericyclic reactions for organic synthesis; many thousands of examples of all three are known. Note that the Claisen rearrangement works well with both allyl aryl ethers and with allyl vinylic ethers.

Claisen rearrangement



Claisen rearrangement

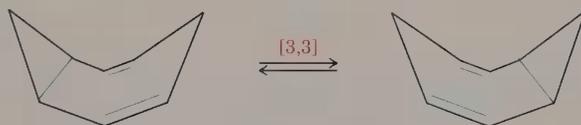


Cope rearrangement



Since both Cope and Claisen rearrangements involve reorganization of an odd number of electron pairs (two pi bonds and one sigma bond), they are predicted to react by suprafacial pathways (Figure 30.21 on page 1176).

The ease with which these two rearrangements occur provides strong evidence for the correctness of orbital-symmetry predictions. For example, it has been estimated, based on spectroscopic measurements, that homotropilidene undergoes [3,3] Cope rearrangement *several hundred times each second* at room temperature.



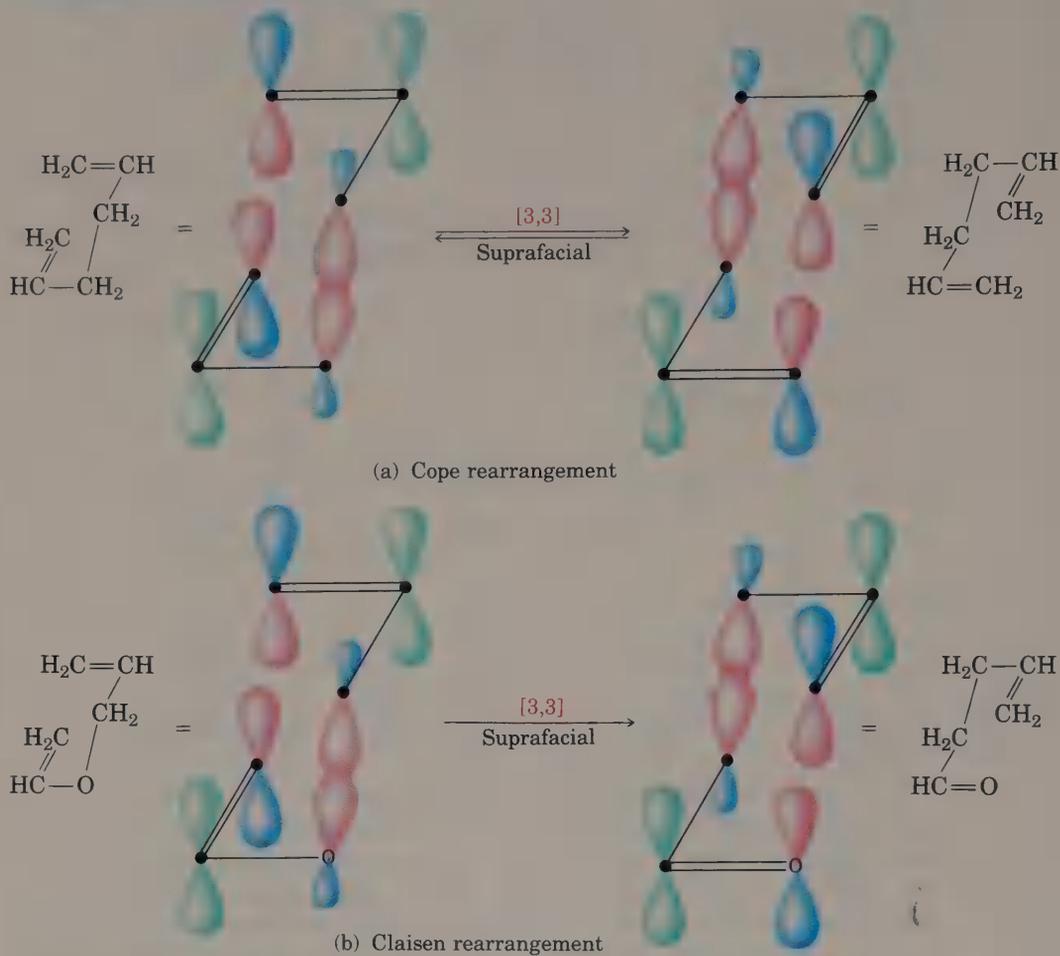
**Homotropilidene**

PROBLEM .....

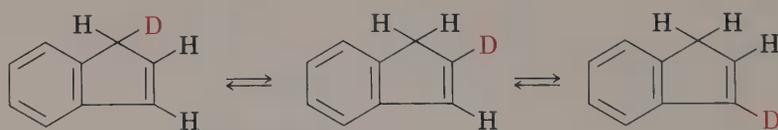
- 30.10 The  $^{13}\text{C}$  NMR spectrum of homotropilidene taken at room temperature shows only three peaks. Explain.

PROBLEM .....

- 30.11 How can you account for the fact that heating 1-deuterioindene scrambles the isotope label to all three positions on the five-membered ring?

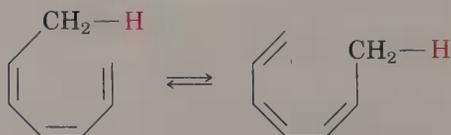


**Figure 30.21** Suprafacial [3,3] Cope and Claisen rearrangements.



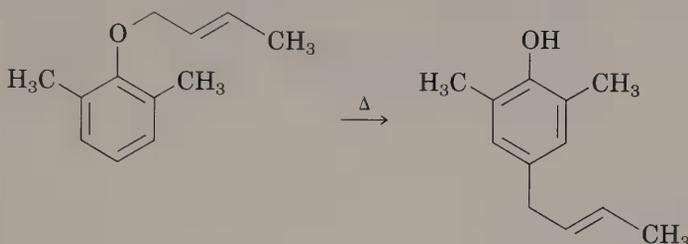
**PROBLEM** .....

**30.12** Classify the following sigmatropic reaction by order  $[x,y]$ , and indicate whether you would expect it to proceed with suprafacial or antarafacial stereochemistry.



## PROBLEM

- 30.13 When a 2,6-disubstituted allyl phenyl ether is heated in an attempted Claisen rearrangement, migration occurs to give *p*-allyl product. Explain how this occurs.



### 30.13 A Summary of Rules for Pericyclic Reactions

Pericyclic, electrocyclic, cycloaddition, sigmatropic, conrotatory, disrotatory, suprafacial, antarafacial . . . how can you keep it all straight?

The information provided in Tables 30.1, 30.2, and 30.3 to summarize the selection rules for electrocyclic, cycloaddition, and sigmatropic reactions leads to the conclusion that pericyclic processes can be grouped according to whether they involve the reorganization of an even or an odd number of electron pairs (bonds). All of this information can be distilled into one simple phrase that provides an easy and accurate way to predict the stereochemical outcome of any pericyclic reaction:

For a ground-state (thermal) pericyclic reaction, the groupings are *odd-supra-dis* and *even-antara-con*.

Cycloaddition and sigmatropic reactions involving an odd number of electron pairs (bonds) occur with suprafacial geometry; electrocyclic reactions involving an odd number of electron pairs occur with disrotatory stereochemistry. Conversely, pericyclic reactions involving an even number of electron pairs occur with either antarafacial geometry or conrotatory stereochemistry.

Once the selection rules for thermal reactions have been memorized, the rules for photochemical reactions are derived simply by remembering that they're the opposite of the thermal rules:

For an excited-state (photochemical) pericyclic reaction, the groupings are *odd-antara-con* and *even-supra-dis*.

Both rules are summarized in Table 30.4. Memorizing this table will give you the ability to predict the stereochemistry of literally thousands of pericyclic reactions.

TABLE 30.4 Generalized Selection Rules for Pericyclic Reactions

<i>Electron state</i>	<i>Electron pairs</i>	<i>Stereochemistry</i>
Ground state (thermal)	Even number	Antara-con
	Odd number	Supra-dis
Excited state (photochemical)	Even number	Supra-dis
	Odd number	Antara-con

## PROBLEM.....

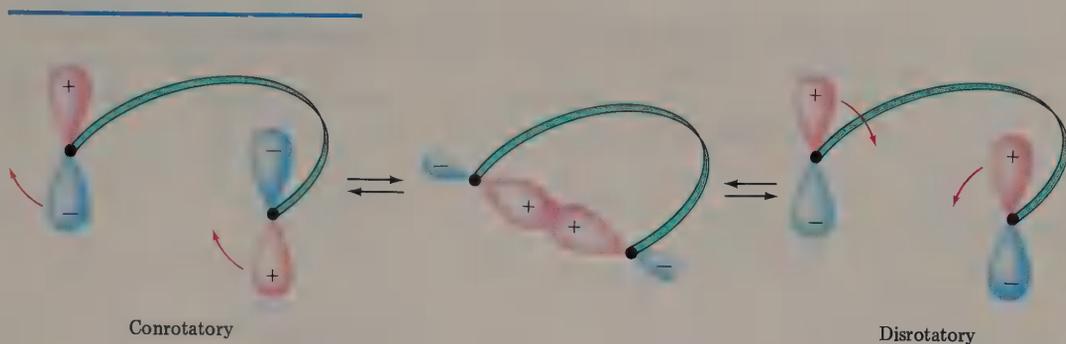
- 30.14 Predict the stereochemistry of these pericyclic reactions:
- The thermal cyclization of a conjugated tetraene
  - The photochemical cyclization of a conjugated tetraene
  - A photochemical [4 + 4] cycloaddition
  - A thermal [2 + 6] cycloaddition
  - A photochemical [3,5] sigmatropic rearrangement
- .....

### 30.14 Summary and Key Words

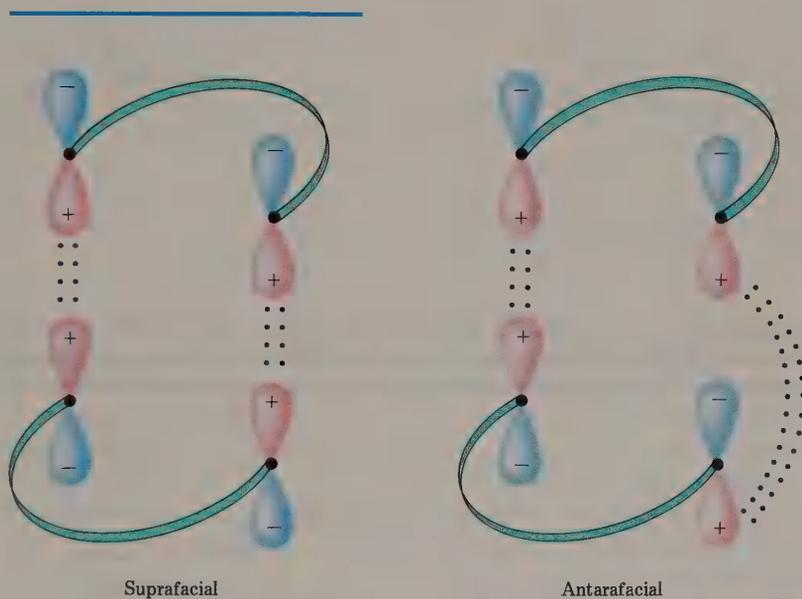
A **pericyclic reaction** is one that takes place in a single step through a cyclic transition state; no intermediates are involved. There are three major classes of pericyclic processes: **electrocyclic reactions**, **cycloaddition reactions**, and **sigmatropic rearrangements**. The stereochemistry of these reactions is controlled by the symmetry of the orbitals involved in bond reorganization.

Electrocyclic reactions involve the cyclization of conjugated polyenes. For example, 1,3,5-hexatriene cyclizes to 1,3-cyclohexadiene on heating. Electrocyclic reactions can occur by either **conrotatory** or **disrotatory** paths, depending on the symmetry of the terminal lobes of the pi system. Conrotatory cyclization requires that both lobes rotate in the same direction, whereas disrotatory cyclization requires that the lobes rotate in opposite directions (Figure 30.22). The reaction course for any specific case can be found by looking at the symmetry of the **highest occupied molecular orbital (HOMO)**.

Cycloaddition reactions are those in which two unsaturated molecules add together to yield a cyclic product. For example, Diels–Alder reaction between a diene (4 pi electrons) and a dienophile (2 pi electrons) yields a cyclohexene. Cycloadditions can take place by either **suprafacial** or **antarafacial** pathways. Suprafacial cycloaddition involves reaction between lobes on the same face of one component and on the same face of the second component. Antarafacial cycloaddition involves reaction between lobes on the same face of one component and on opposite faces of the other component (Figure 30.23). The reaction course in a specific case can be found by looking at the symmetry of the HOMO of one component and the LUMO of the other component.



**Figure 30.22** Conrotatory and disrotatory motions during electrocyclic reactions.



**Figure 30.23** Suprafacial and antarafacial cycloadditions.

Sigmatropic rearrangements involve the migration of a sigma-bonded group across a pi electron system. For example, **Claisen rearrangement** of an allyl vinyl ether yields an unsaturated carbonyl compound, and **Cope rearrangement** of a 1,5-hexadiene yields a new 1,5-hexadiene. Sigmatropic rearrangements can occur with either suprafacial or antarafacial stereochemistry; the selection rules for a given case are the same as those for cycloaddition reactions.

The stereochemistry of any pericyclic reaction can be predicted by counting the total number of electron pairs (bonds) involved in bond reorganization and then applying some simple rules. Thermal (ground-state) reactions involving an even number of electron pairs occur with either

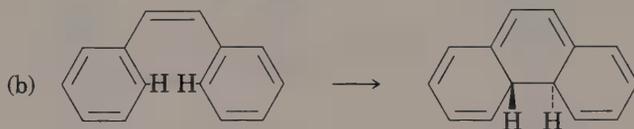
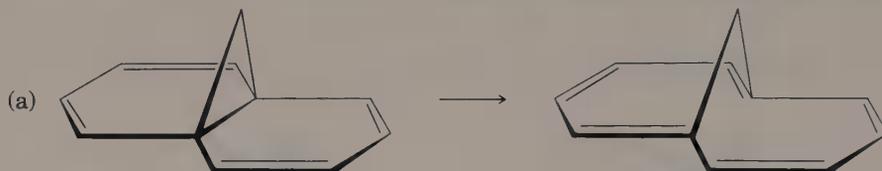
antarafacial or conrotatory stereochemistry (**even-antara-con**); thermal reactions involving an odd number of electron pairs occur with suprafacial or disrotatory stereochemistry (**odd-supra-dis**). Exactly the opposite rules apply to photochemical (excited-state) reactions.

### ADDITIONAL PROBLEMS .....

30.15 Define these terms in your own words:

- |                            |                               |
|----------------------------|-------------------------------|
| (a) Electrocyclic reaction | (b) Conrotatory motion        |
| (c) Suprafacial            | (d) Antarafacial              |
| (e) Disrotatory motion     | (f) Sigmatropic rearrangement |

30.16 Have the following reactions taken place in a conrotatory or in a disrotatory manner? Under what conditions, thermal or photochemical, would you carry out each reaction?



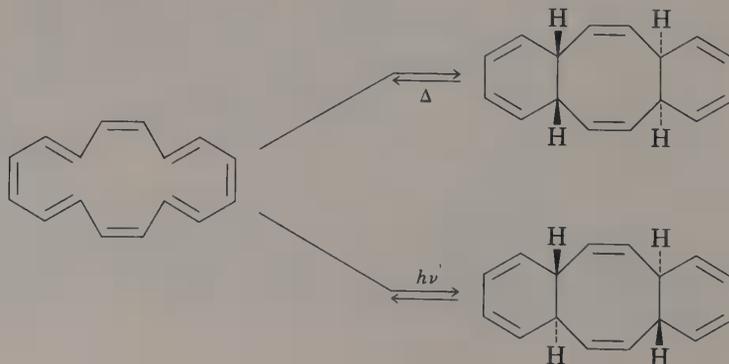
30.17 (*2E,4Z,6Z,8E*)-Decatetraene has been cyclized to give 7,8-dimethyl-1,3,5-cyclooctatriene. Predict the manner of ring closure—conrotatory or disrotatory—for both thermal and photochemical reactions, and predict the stereochemistry of the product in each case.

30.18 Answer Problem 30.17 for the thermal and photochemical cyclizations of (*2E,4Z,6Z,8Z*)-decatetraene.

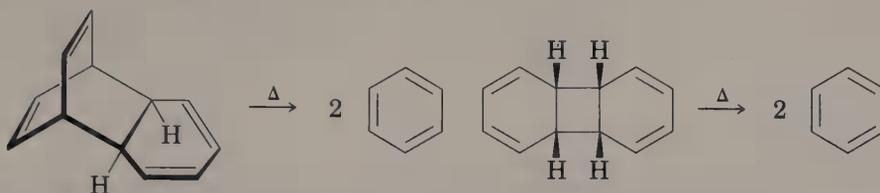
30.19 What stereochemistry would you expect to observe in these reactions?

- A photochemical [1,5] sigmatropic rearrangement
- A thermal [4 + 6] cycloaddition
- A thermal [1,7] sigmatropic rearrangement
- A photochemical [2 + 6] cycloaddition

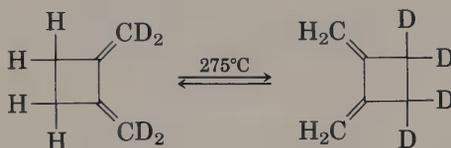
30.20 The cyclohexadecaoctaene shown isomerizes to two different isomers, depending on reaction conditions. Explain the observed results, and indicate whether each reaction is conrotatory or disrotatory.



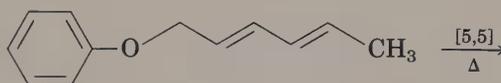
30.21 Which of the following two reactions is more likely to occur? Explain.



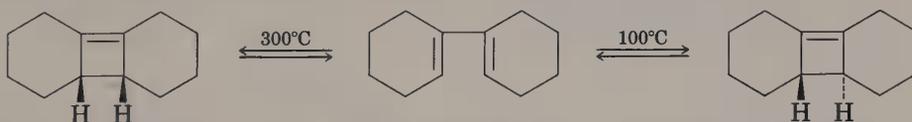
30.22 The following thermal rearrangement involves two pericyclic reactions in sequence. Identify them and propose a mechanism to account for the observed result.



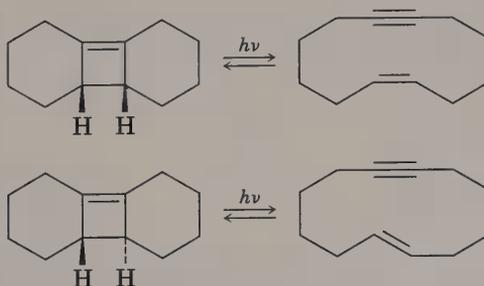
30.23 Predict the product of the following pericyclic reaction. Should this [5,5] shift be a suprafacial or an antarafacial process?



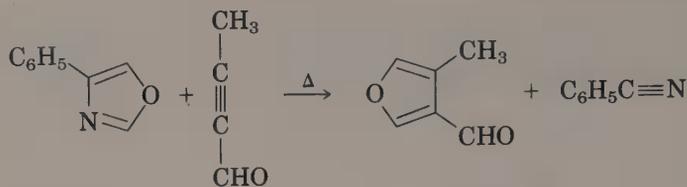
30.24 How can you account for the fact that ring opening of the *trans*-cyclobutene isomer shown takes place at much lower temperatures than a similar ring opening of the *cis*-cyclobutene isomer does? Identify the stereochemistry of each reaction as either conrotatory or disrotatory.



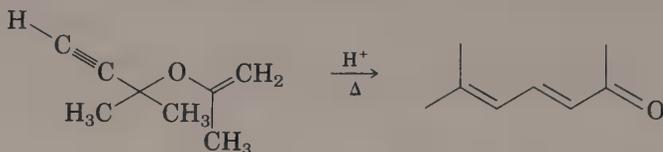
30.25 Photolysis of the *cis*-cyclobutene isomer in Problem 30.24 yields *cis*-cyclododecaen-7-yne, but photolysis of the *trans* isomer yields *trans*-cyclododecaen-7-yne. Explain these results, and identify the type and stereochemistry of the pericyclic reaction.



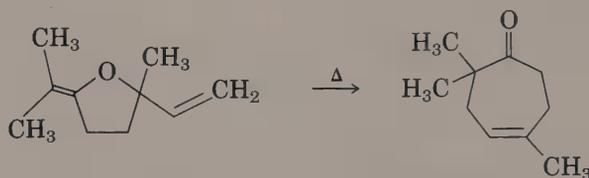
30.26 Two pericyclic reactions are involved in the furan synthesis shown. Identify them and propose a mechanism for the transformation.



- 30.27 The following synthesis of dienones occurs readily. Propose a mechanism to account for the results, and identify the kind of pericyclic reaction involved.



- 30.28 Karahanaenone, a compound isolated from oil of hops, has been synthesized by the thermal reaction shown. Identify the kind of pericyclic reaction, and explain how karahanaenone is formed.



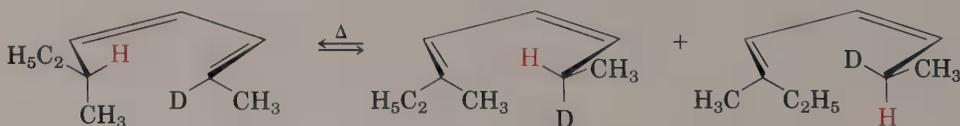
Karahanaenone

- 30.29 The <sup>1</sup>H NMR spectrum of bullvalene at 100°C consists only of a single peak at 4.22 δ. What conclusion can you draw from this? What explanation can you suggest to account for this result?

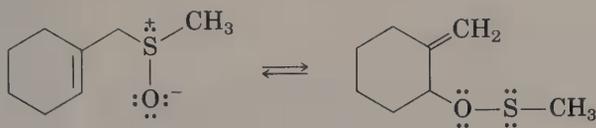


Bullvalene

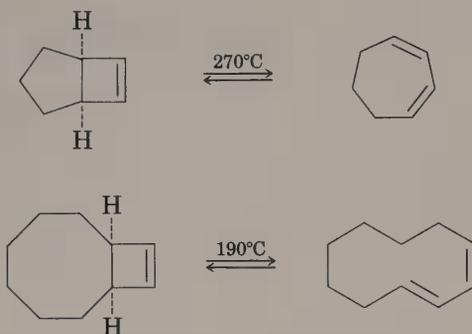
- 30.30 The rearrangement shown was devised and carried out to prove the stereochemistry of [1,5] sigmatropic hydrogen shifts. Explain how the observed result confirms the predictions of orbital symmetry.



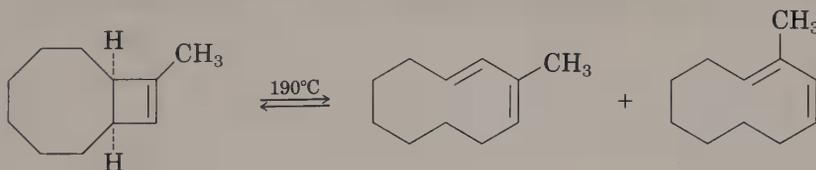
- 30.31 The reaction shown is an example of a [2,3] sigmatropic rearrangement. Would you expect the reaction to be suprafacial or antarafacial? Explain.



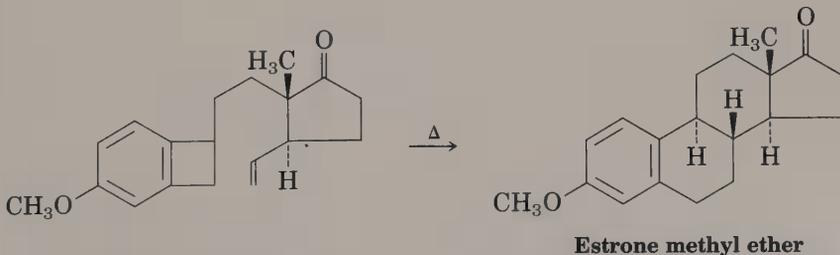
- 30.32 When the compound having a cyclobutene fused to a five-membered ring is heated, (1*Z*,3*Z*)-cycloheptadiene is formed. When the related compound having a cyclobutene fused to an eight-membered ring is heated, however, (1*E*,3*Z*)-cyclodecadiene is formed. Explain these results, and suggest a reason why the eight-membered-ring opening occurs at lower reaction temperatures.



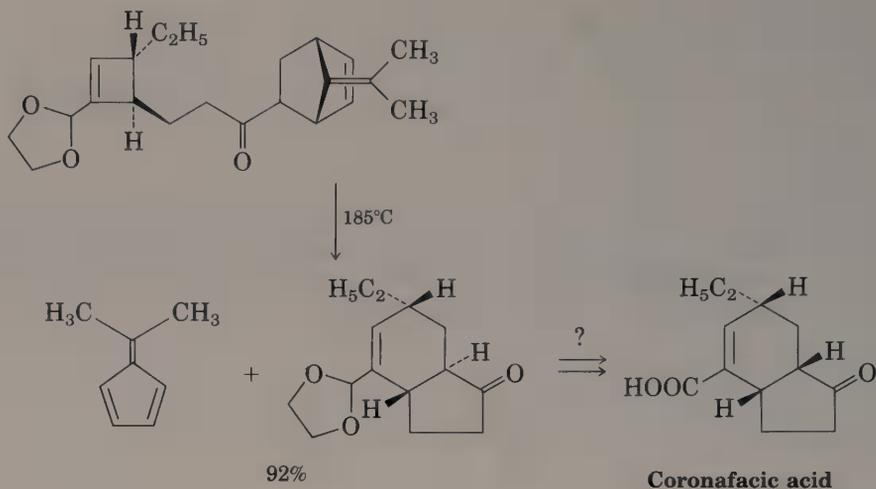
- 30.33 In light of your answer to Problem 30.32, explain why a mixture of products occurs in the following reaction:



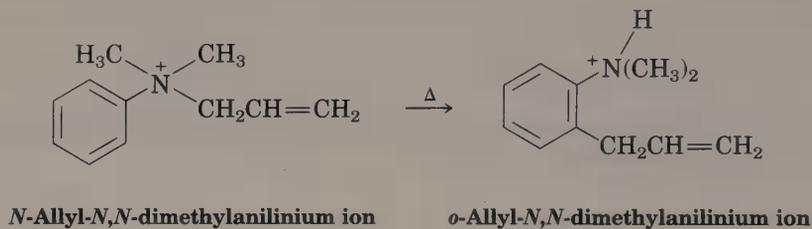
- 30.34 Estrone, a major female sex hormone, has been synthesized by a route that involves the following step. Identify the pericyclic reactions involved, and propose a mechanism.

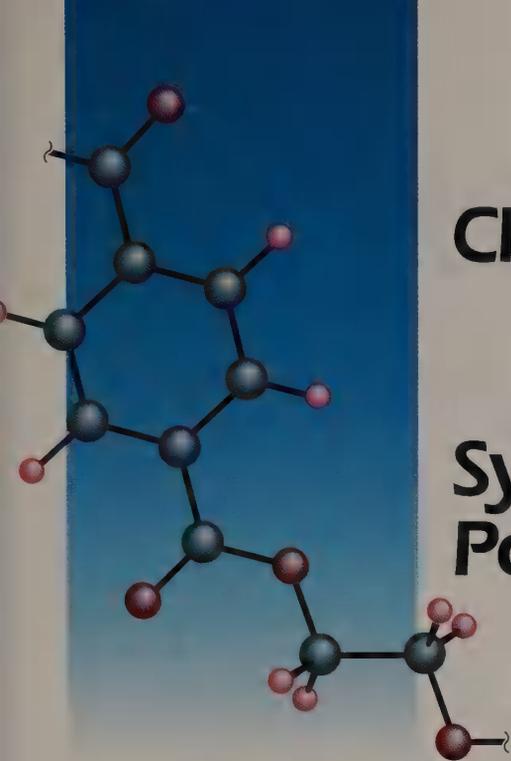


- 30.35 Coronafacic acid, a bacterial toxin, was synthesized using a key step that involves three sequential pericyclic reactions. Identify them and propose a mechanism for the overall transformation. How would you complete the synthesis?



**30.36** The following rearrangement of *N*-allyl-*N,N*-dimethylanilinium ion has been observed. Propose a mechanism to account for the reaction.





# Chapter 31

## Synthetic Polymers

No other group of compounds has had as great an impact on our day-to-day living as the synthetic polymers. As plastics, adhesives, and paints, synthetic polymers have a multitude of uses, from the foam coffee cup to the life-saving artificial heart valve. Polymer synthesis is a major part of the chemical industry, and annual production figures for some of the more important polymers are shown in Table 31.1.

**TABLE 31.1** Production Figures for Some Major Polymers (1990)

<i>Polymer</i>	<i>U.S. production (million ton/yr)</i>	<i>Polymer</i>	<i>U.S. production (million ton/yr)</i>
Polyethylene		Acrylonitrile/butadiene/styrene	1.17
Low-density	5.59	Acrylic fibers	0.25
High-density	4.17	Epoxy adhesives	0.25
Polypropylene	4.16	Phenolic resins	1.47
Polystyrene	2.51	Urea/formaldehyde resins	0.75
Poly(vinyl chloride)	4.54	Styrene/butadiene rubber	0.43
Nylon	1.33	Polybutadiene rubber	0.20
Polyesters	1.60		

A **polymer** is simply a large molecule built up by repetitive bonding together of many smaller units called **monomers**. We've already seen, for example, that cellulose is a large carbohydrate polymer built of repeating glucose units, proteins are large polyamides built of repeating amino acid units, and nucleic acids are large molecules built of repeating nucleotide units.

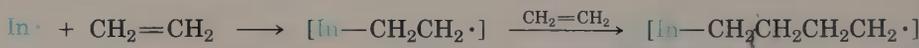
Synthetic polymers are chemically much simpler than most biopolymers, since the repeating monomer units tend to be small, simple, and inexpensive molecules. There is, however, an immense diversity to the structure and properties of synthetic polymers, depending on the structure of the monomer and on the reaction conditions used for polymerization.

Polymers can be classified in many ways—by method of synthesis, by structure, by physical properties, or by end use, to name a few. Let's begin by looking at the organic chemistry of polymer synthesis and then see how polymer structure is correlated with uses and physical properties.

### 31.1 General Classes of Polymers

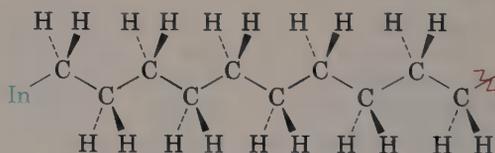
Synthetic polymers can be classified by their method of synthesis as either chain-growth polymers or step-growth polymers. The categories are necessarily broad and imprecise but nevertheless provide a useful distinction. **Chain-growth polymers**, also called *addition polymers*, are produced by chain-reaction polymerization in which an initiator adds to a carbon-carbon double bond to yield a reactive intermediate. This intermediate reacts with a second molecule of monomer to yield a new intermediate, this new intermediate reacts with a third monomer unit, and so on. The polymer is built as more monomers add successively to the reactive end of the growing chain.

The initiator can be an anion, a cation, or a radical, and the monomer unit can be any of a large number of substituted alkenes. Polyethylene, produced by radical-initiated polymerization of ethylene, is by far the most common example of a chain-growth polymer.



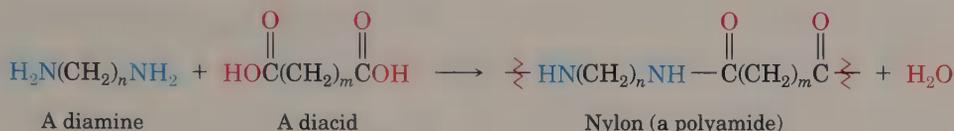
A radical  
initiator

↓ Repeat  
many times



A section of a polyethylene chain

**Step-growth polymers**, also called *condensation polymers*, are produced by processes in which the bond-forming step is one of the fundamental polar reactions we've studied. Reactions occur between two difunctional molecules, with each bond in the polymer being formed independently of the others. The polymer produced normally has the two monomers in an alternating order and usually has other atoms in addition to carbon in the main chain. Nylon, a polyamide formed by reaction of a diacid and a diamine, is the most common example of a step-growth polymer:

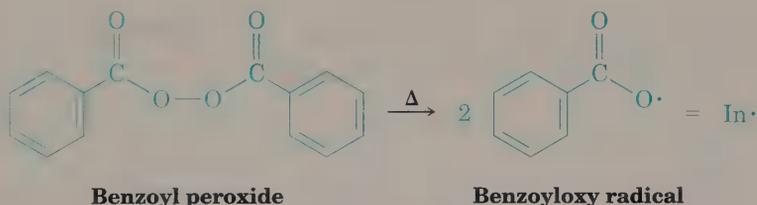


## 31.2 Radical Polymerization of Alkenes

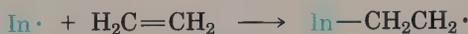
Many low-molecular-weight alkenes undergo a rapid polymerization reaction when treated with small amounts of a radical initiator. For example, polyethylene, one of the first alkene chain-growth polymers to be manufactured commercially, has been produced since 1943 and has a current annual U.S. production volume of more than 9 million tons.

Polymerization of ethylene is usually carried out at high pressure (1000–3000 atm) and high temperature (100–250°C) with a radical catalyst such as benzoyl peroxide. The resultant polymer may have anywhere from a few hundred to a few thousand monomer units in the chain. As with all radical chain reactions, three kinds of steps are required: initiation steps, propagation steps, and termination steps.

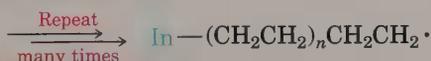
*Step 1* *Initiation* occurs when trace amounts of radicals are generated by the catalyst:



*Step 2* One of the benzoyloxy radicals produced in step 1 adds to ethylene to generate a new carbon radical, and the polymerization is off and running:



*Step 3* *Propagation* occurs when the carbon radical adds to another ethylene molecule. Repetition of this step builds the polymer chain.



*Step 4* Eventually, the polymer chain is *terminated* by reactions that consume the radical. Combination or disproportionation of two radicals are possible chain-terminating reactions:



Many substituted ethylene monomers undergo radical chain polymerization to yield polymers with substituent groups (denoted by a circled S) regularly spaced along the polymer backbone.

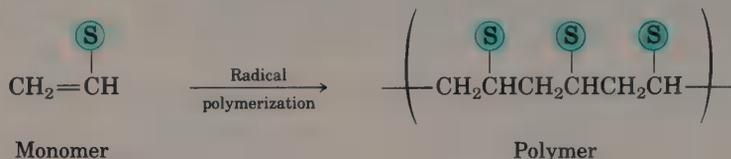
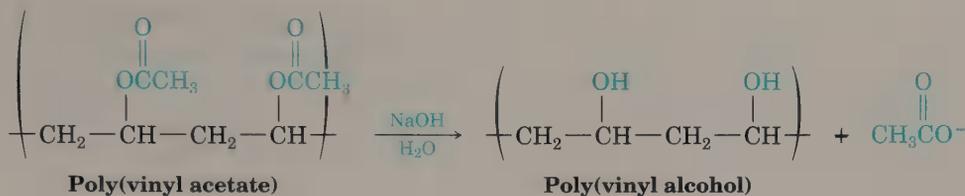


Table 31.2 shows some of the more important of these **vinyl monomers** and lists the industrial uses of the different polymers that result.

**TABLE 31.2** Some Chain-Growth Polymers and Their Uses

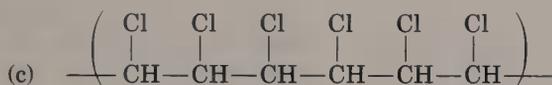
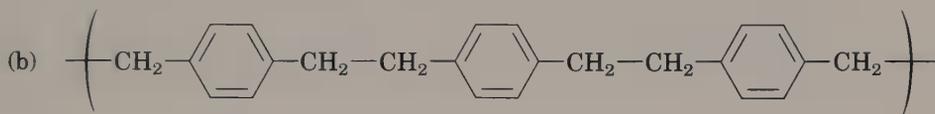
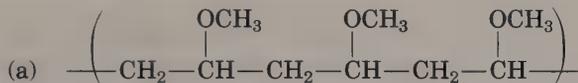
<i>Monomer name</i>	<i>Formula</i>	<i>Trade or common name of polymer</i>	<i>Uses</i>
Ethylene	$\text{H}_2\text{C}=\text{CH}_2$	Polyethylene	Packaging, bottles, cable insulation, films and sheets
Propene (propylene)	$\text{H}_2\text{C}=\text{CHCH}_3$	Polypropylene	Automotive moldings, rope, carpet fibers
Chloroethylene (vinyl chloride)	$\text{H}_2\text{C}=\text{CHCl}$	Poly(vinyl chloride), Tedlar	Insulation, films, pipes
Styrene	$\text{H}_2\text{C}=\text{CHC}_6\text{H}_5$	Polystyrene, Styron	Foam and molded articles
Tetrafluoroethylene	$\text{F}_2\text{C}=\text{CF}_2$	Teflon	Valves and gaskets, coatings
Acrylonitrile	$\text{H}_2\text{C}=\text{CHCN}$	Orlon, Acrilan	Fibers
Methyl methacrylate	$\text{H}_2\text{C}=\overset{\text{CH}_3}{\underset{ }{\text{C}}}\text{CO}_2\text{CH}_3$	Plexiglas, Lucite	Molded articles, paints
Vinyl acetate	$\text{H}_2\text{C}=\text{CHOCOCH}_3$	Poly(vinyl acetate)	Paints, adhesives
Vinyl alcohol	" $\text{H}_2\text{C}=\text{CHOH}$ "	Poly(vinyl alcohol)	Fibers, adhesives

Note that vinyl alcohol, the monomer corresponding to poly(vinyl alcohol), is an enol that would, if prepared, tautomerize rapidly to acetaldehyde. Poly(vinyl alcohol) is therefore made by polymerization of vinyl acetate ( $\text{H}_2\text{C}=\text{CHOCOCH}_3$ ), followed by hydrolysis of the ester groups in poly(vinyl acetate):



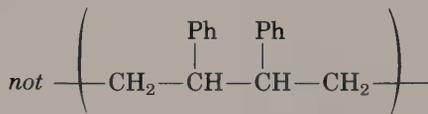
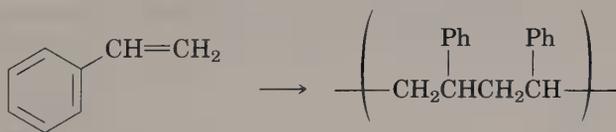
PROBLEM.....

31.1 How might these polymers be prepared? Show the monomer units you would use.



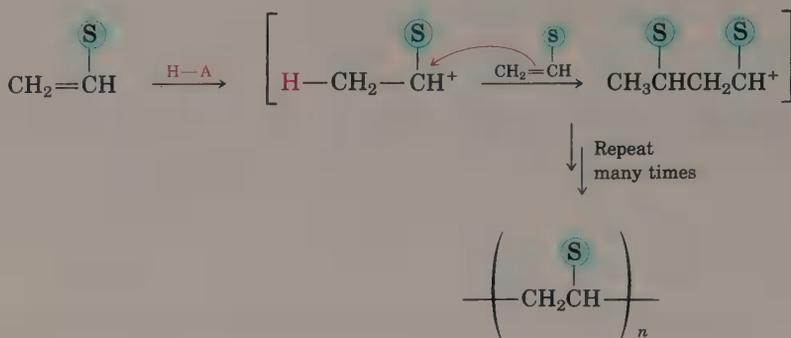
PROBLEM.....

31.2 How can you account for the fact that radical polymerization of styrene yields a product in which the phenyl substituents are on alternate carbon atoms rather than on neighboring carbons?

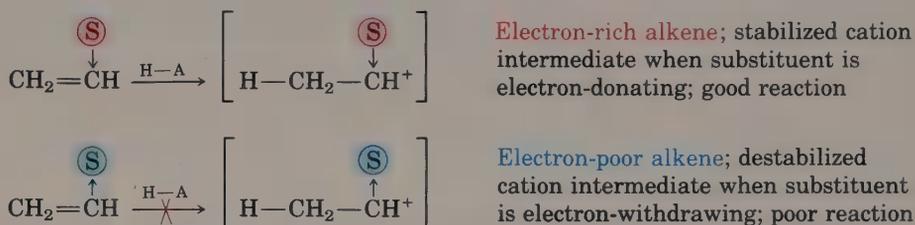


## 31.3 Cationic Polymerization

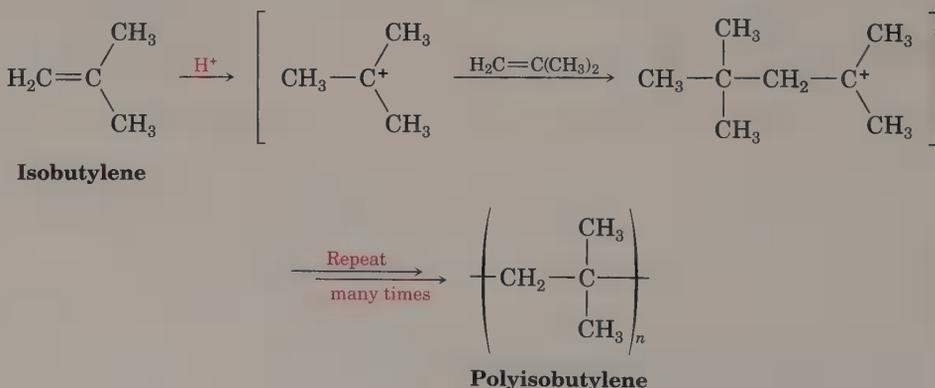
Certain alkene monomers can be polymerized by a *cationic* initiator, as well as by a radical initiator. Cationic polymerizations occur by a chain-reaction pathway and require the use of strong protic or Lewis acids. The key chain-carrying step is the electrophilic addition of a carbocation intermediate to the carbon-carbon double bond of another monomer unit.



As you might expect, vinyl monomers with electron-donating substituents polymerize much more readily than monomers with electron-withdrawing substituents. Thus, ethylene, vinyl chloride, and acrylonitrile don't polymerize easily under cationic conditions, but 2-methylpropene polymerizes rapidly. This difference in behavior simply reflects the difference in stability of the potential chain-carrying intermediate cations:

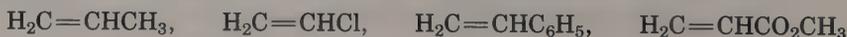


The most common commercial use of cationic polymerization occurs in the preparation of polyisobutylene by treatment of isobutylene (2-methylpropene) with  $\text{BF}_3$  catalyst at  $-80^\circ\text{C}$ . The product is used in the manufacture of inner tubes for truck and bicycle tires.



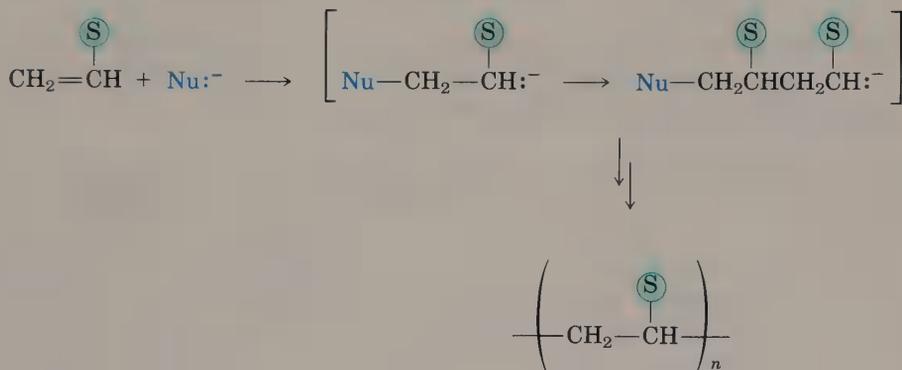
**PROBLEM** .....

**31.3** List the expected reactivity order of the following monomers to cationic polymerization. Explain your ordering.



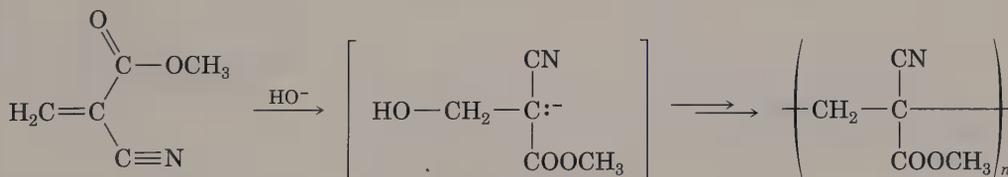
## 31.4 Anionic Polymerization

Alkene monomers with electron-withdrawing, anion-stabilizing substituents can be polymerized by anionic catalysts. A chain reaction occurs in which the key step is nucleophilic addition of an anion to the unsaturated monomer by a Michael reaction (Section 23.12). Acrylonitrile ( $\text{H}_2\text{C}=\text{CHCN}$ ), methyl methacrylate [ $\text{H}_2\text{C}=\text{C}(\text{CH}_3)\text{COOCH}_3$ ], and styrene ( $\text{H}_2\text{C}=\text{CHC}_6\text{H}_5$ ) can all be polymerized anionically, although radical-initiated polymerization is preferred commercially.



where Nu = A nucleophilic initiator  
S = An electron-withdrawing substituent

One particularly interesting example of anionic polymerization accounts for the remarkable properties of “super glue,” one drop of which can support up to 2000 lb. Super glue is simply a solution of pure methyl  $\alpha$ -cyanoacrylate. Since the carbon-carbon double bond has two electron-withdrawing groups, anionic addition is particularly easy. Trace amounts of water or bases on the surface of an object are sufficient to initiate polymerization of the cyanoacrylate and bind articles together. Skin is a good source of the necessary basic initiators, and many people have found their fingers stuck together after inadvertently touching super glue.

Methyl  $\alpha$ -cyanoacrylate

Super glue

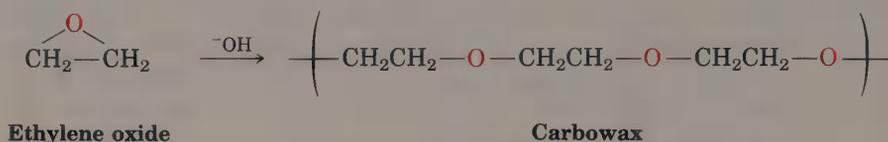
PROBLEM.....

31.4 Order the following monomers with respect to their expected reactivity toward anionic polymerization. Explain your ordering.



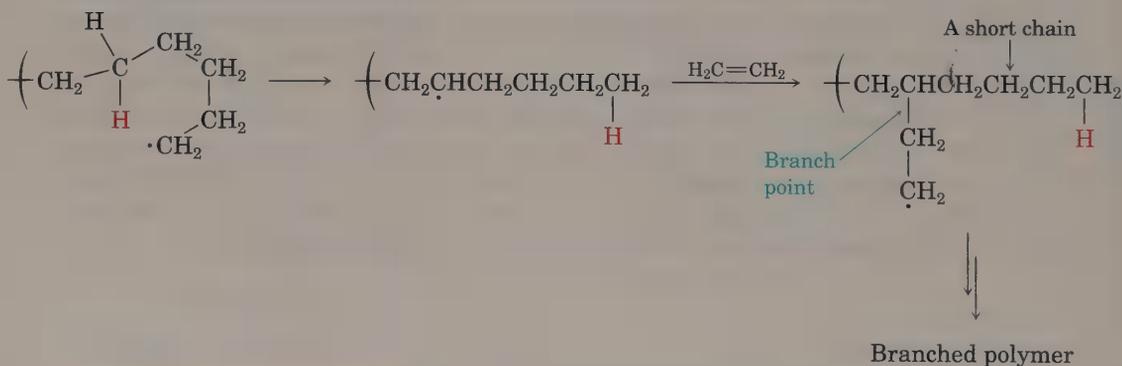
## PROBLEM.....

- 31.5 Poly(ethylene glycol), or Carbowax, can be made by base-induced polymerization of ethylene oxide. Propose a mechanism for this reaction.



### 31.5 Chain Branching during Polymerization

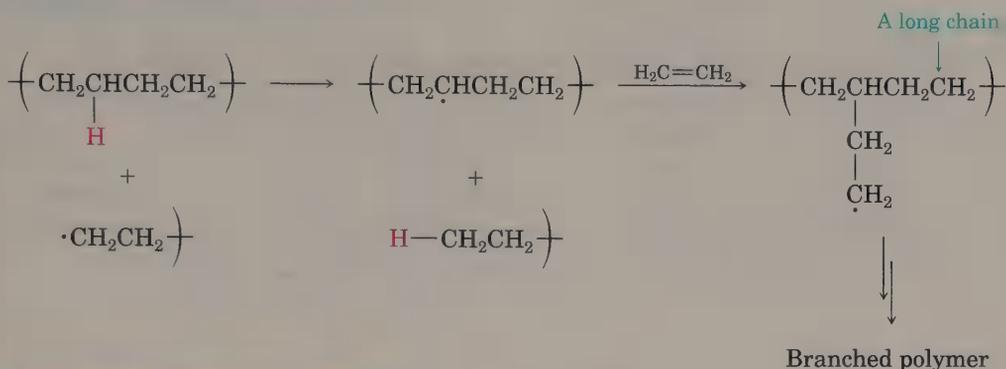
The polymerization of unsaturated monomers is complicated in practice by several problems that greatly affect the properties of the product. One such problem is that radical polymerization yields a product that is not linear but has numerous branches in it. Branches arise when the radical end of a growing chain abstracts a hydrogen atom from the middle of the chain to yield an internal radical site that continues the polymerization. The most common kind of branching, termed **short-chain branching**, arises from intramolecular hydrogen atom abstraction from a position four carbon atoms away from the chain end (Figure 31.1).



**Figure 31.1** Short-chain branching during the polymerization of ethylene.

Alternatively, intermolecular hydrogen atom abstraction can take place by reaction of the radical end of one chain with the middle of another chain. **Long-chain branching** results from this kind of reaction (Figure 31.2). Studies have shown that short-chain branching occurs about 50 times more often than long-chain branching.

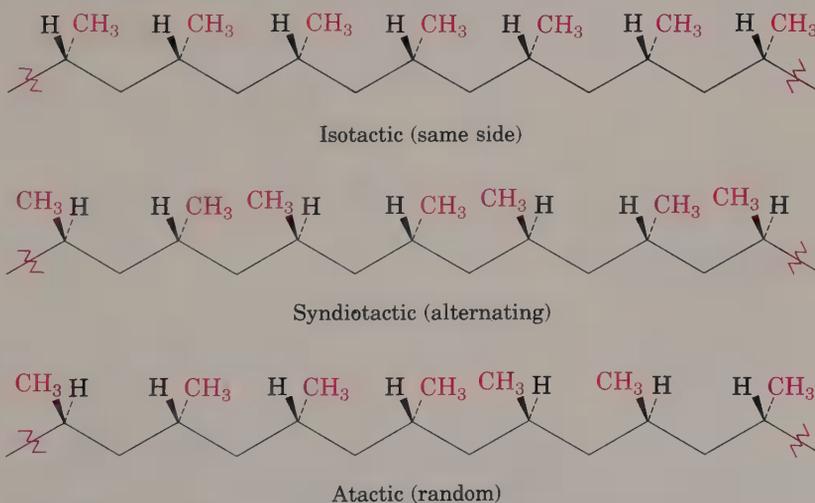
Chain branching is a common occurrence during radical polymerizations and is not restricted to polyethylene. Polypropylene, polystyrene, and poly(methyl methacrylate) all contain branched chains.



**Figure 31.2** Long-chain branching during the polymerization of ethylene.

## 31.6 Stereochemistry of Polymerization: Ziegler–Natta Catalysts

Another complication that arises during alkene polymerization has to do with stereochemistry. Although not pointed out earlier, polymerization of substituted alkenes leads to polymers with numerous stereogenic centers on the backbone. For example, propylene might polymerize with any of the three stereochemical outcomes shown in Figure 31.3. The conformation in which all methyl groups are on the same side of the zigzag backbone is called **isotactic**; that in which the methyl groups regularly alternate on opposite sides of the backbone is called **syndiotactic** (*syn-di-o-tac-tic*); and the conformation in which the methyl groups are randomly oriented is called **atactic**.



**Figure 31.3** Isotactic, syndiotactic, and atactic forms of polypropylene.

The three different stereochemical forms of polypropylene all have somewhat different properties, and all three can be made by the proper choice of polymerization conditions. Branched atactic polymers arise from normal radical chain polymerizations, but the use of special **Ziegler–Natta**<sup>1,2</sup> **catalysts** allows preparation of isotactic and syndiotactic forms.

Ziegler–Natta catalysts are organometallic transition-metal complexes prepared by treatment of a trialkylaluminum with a titanium compound. Triethylaluminum and titanium trichloride form a typical preparation, although the precise structure of the active catalyst is still unknown.

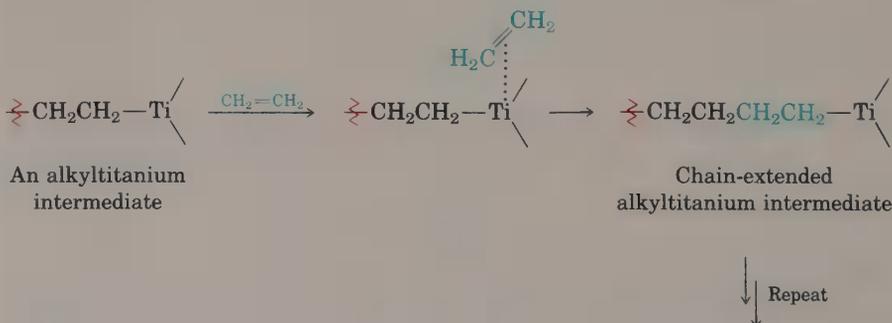


Introduced in 1953, Ziegler–Natta catalysts immediately revolutionized the field of polymer chemistry, largely because of two advantages:

1. Ziegler–Natta polymers are linear and have practically no chain branching.
2. Ziegler–Natta polymers are stereochemically regular. Either isotactic or syndiotactic forms can be produced, depending on the exact catalyst system used. All commercial polypropylene is now produced by the Ziegler–Natta process, since the product has greater strength, stiffness, and resistance to cracking than does the branched atactic polypropylene prepared by radical polymerization.

Linear polyethylene produced by the Ziegler–Natta process (called *high-density polyethylene*) is a highly crystalline polymer with 500–1000 ethylene units per chain. High-density polyethylene has greater strength and heat resistance than the product of radical-induced polymerization (low-density polyethylene) has and is used to produce plastic squeeze bottles and molded housewares.

The exact mechanism by which Ziegler–Natta catalysts operate is unclear, but the key chain-lengthening steps undoubtedly involve formation of alkyltitanium species, followed by coordination of alkene monomer to the titanium and insertion of coordinated alkene into the carbon–titanium bond.



<sup>1</sup>Karl Ziegler (1889–1976); b. Helsa, near Kassel, Germany; Ph.D. Marburg University; director, Max Planck Institute for Coal Research, Mülheim-Ruhr, Germany; Nobel Prize (1963).

<sup>2</sup>Giulio Natta (1903–1979); b. Imperia, Italy; D.C.E. Milan Polytechnic Institute; professor, Milan Polytechnic Institute; Nobel Prize (1963).

## PROBLEM.....

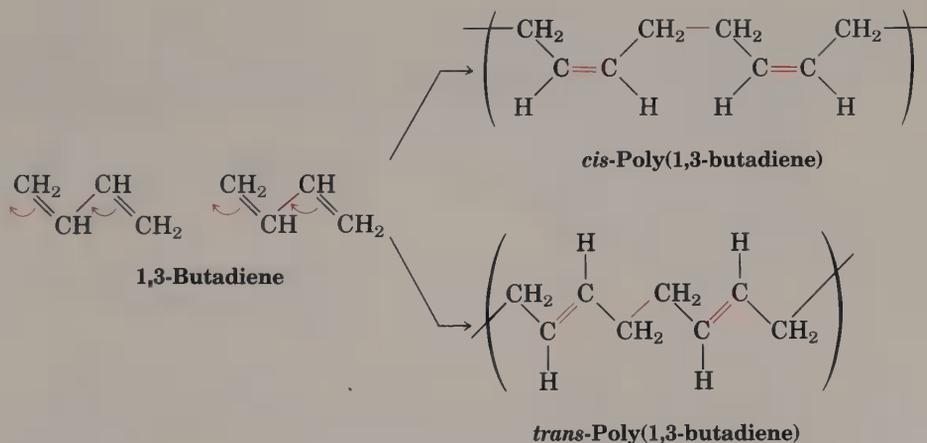
- 31.6 Account for the fact that vinylidene chloride,  $\text{H}_2\text{C}=\text{CCl}_2$ , does not polymerize in isotactic, syndiotactic, and atactic forms.

## PROBLEM.....

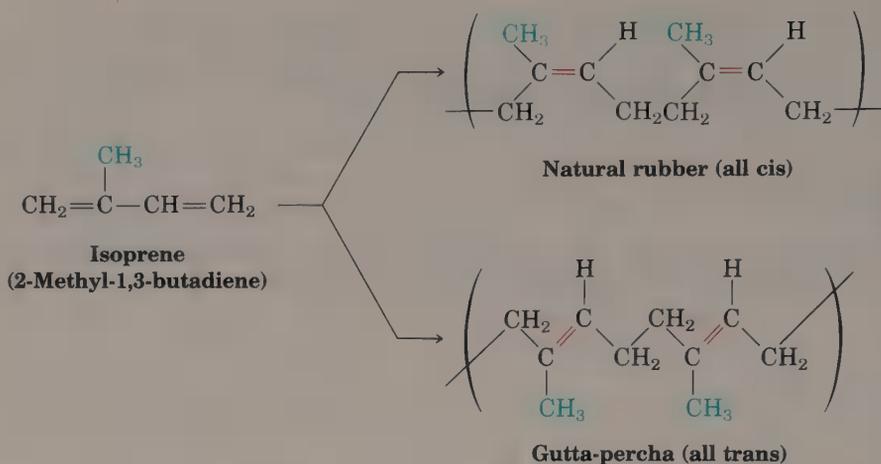
- 31.7 Polymers such as polypropylene contain a large number of chiral carbon atoms. Would you therefore expect samples of isotactic, syndiotactic, or atactic polypropylene to rotate plane-polarized light? Explain.

## 31.7 Diene Polymers: Natural and Synthetic Rubbers

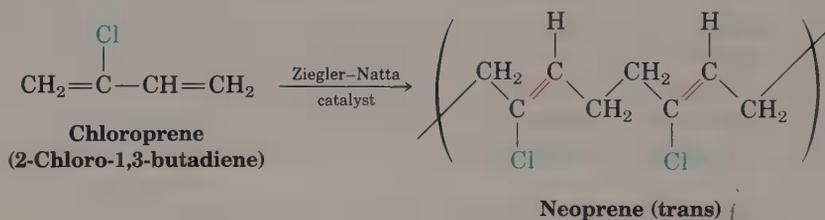
Although we've discussed only the polymerization of simple alkene monomers up to this point, the same principles apply to the polymerization of conjugated dienes. Diene polymers are structurally more complex, though, since double bonds remain every four carbon atoms along the chain. These double bonds can be either *cis* or *trans*, and the proper choice of Ziegler–Natta catalyst allows preparation of either geometry. Note that the polymerization reaction corresponds to 1,4 addition of the growing chain to each conjugated diene monomer (recall 1,4 ionic addition to dienes, Section 14.5).



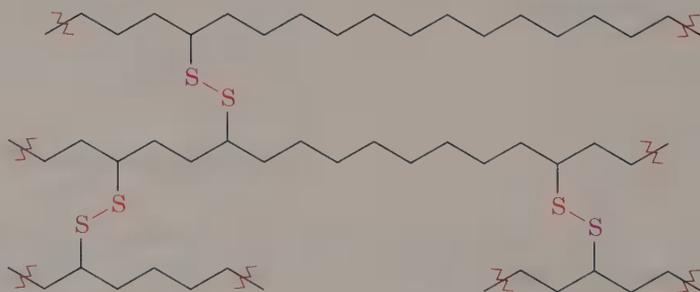
Natural rubber is a polymer of isoprene (2-methyl-1,3-butadiene) in which the double bonds have *cis* stereochemistry. Gutta-percha, the all-*trans* isomer of natural rubber, also occurs naturally as the exudate of certain trees but is less common. Gutta-percha is harder and more brittle than rubber but finds a variety of minor applications, including occasional use as the covering on golf balls.



A number of different synthetic rubbers are produced commercially by diene polymerization. Both *cis*- and *trans*-polyisoprene can be made under Ziegler–Natta conditions, and the synthetic rubber thus produced is similar to the natural material. Chloroprene (2-chloro-1,3-butadiene) is polymerized commercially to yield neoprene, an excellent, though expensive, synthetic rubber with good weather resistance. Neoprene is used in the production of industrial hoses and gloves, among other things.



Both natural and synthetic rubbers are soft and tacky unless hardened by a process called **vulcanization**. Discovered in 1839 by Charles Goodyear (of subsequent tire fame), vulcanization involves heating the polymer with a few percent by weight of sulfur. The result is a much harder rubber with greatly improved resistance to wear and abrasion.

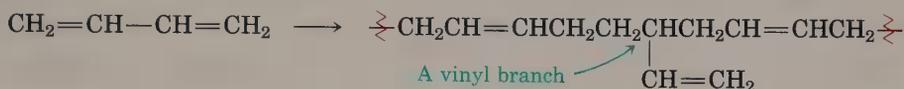


**Figure 31.4** Sulfur cross-linked chains resulting from vulcanization of rubber.

The chemistry of vulcanization involves formation of sulfur bridges or cross-links between polymer chains. Cross-linked polymers tend to be rigid, because the individual chains are locked together into immense single molecules that can no longer slip over one another (Figure 31.4). Note the similarity in structure between a vulcanized rubber and a peptide with cysteine cross-links (Section 27.7).

PROBLEM.....

- 31.8 Diene polymers contain occasional vinyl branches along the chain. How do you think these branches might arise?

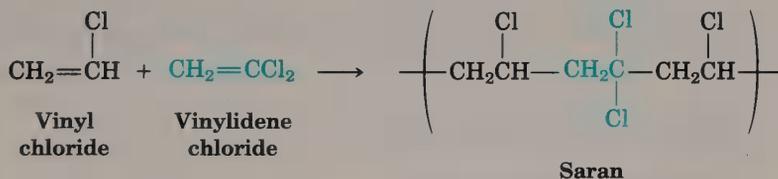


PROBLEM.....

- 31.9 Radial tires, whose sidewalls are made of natural rubber, tend to crack and weather rapidly in areas around cities where high levels of ozone and other industrial pollutants are found. Explain.
- .....

## 31.8 Copolymers

Up to this point we've discussed only **homopolymers**, polymers that are made up of identical repeating units. In practice, however, *copolymers* are more common and more important commercially. **Copolymers** are obtained when two or more different monomers are allowed to polymerize together. For example, copolymerization of vinyl chloride with vinylidene chloride (1,1-dichloroethylene) leads to the well-known polymer Saran.

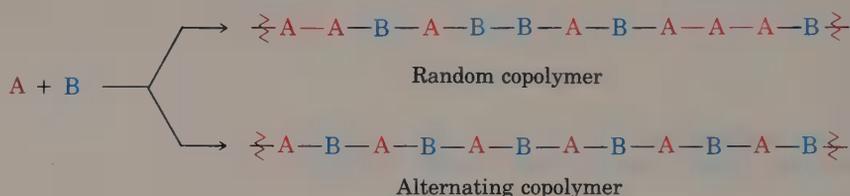


Copolymerization of monomer mixtures often leads to materials with properties quite different from those of either corresponding homopolymer, giving the polymer chemist a vast amount of flexibility for devising new materials. Table 31.3 lists some common copolymers and indicates their commercial applications.

Several different structural types of copolymers can be defined, depending on the distribution of monomer units in the chain. If monomer A and monomer B are being copolymerized, for instance, the resultant product might have a random distribution of the two units throughout the chain, or it might have an alternating distribution:

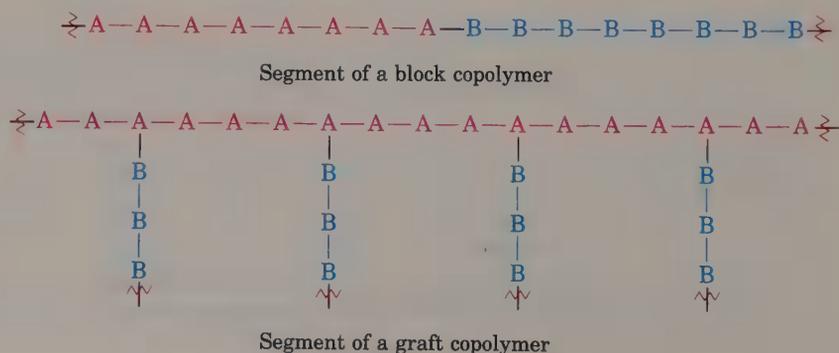
TABLE 31.3 Some Common Copolymers and Their Uses

Monomer name	Formula	Trade or common name of polymer	Uses
Vinyl chloride	$H_2C=CHCl$	Saran	Food wrapping, fibers
Vinylidene chloride	$H_2C=C(Cl)_2$		
Styrene (25%)	$H_2C=CHC_6H_5$	SBR (styrene-butadiene rubber)	Tires
Butadiene (75%)	$H_2C=CHCH=CH_2$		
Hexafluoropropene	$F_2C=CF_2$	Viton	Gaskets, rubber articles
Vinylidene fluoride	$H_2C=CF_2$		
Acrylonitrile	$H_2C=CHCN$	Nitrile rubber	Latex, adhesives, gasoline hoses
Butadiene	$H_2C=CHCH=CH_2$		
Isobutylene	$H_2C=C(CH_3)_2$	Butyl rubber	Inner tubes
Isoprene	$H_2C=C(CH_3)CH=CH_2$		
Acrylonitrile	$H_2C=CHCN$	ABS (initials of three monomers)	Pipes, high-impact applications
Butadiene	$H_2C=CHCH=CH_2$		
Styrene	$H_2C=CHC_6H_5$		



The exact distribution depends on such factors as the initial proportion of the two reactant monomers and their relative reactivities. In practice, neither perfectly random nor perfectly alternating copolymers are usually found. Most copolymers tend more toward the alternating form but have many random imperfections.

Two other special forms of copolymers that can be prepared under certain conditions are called *block copolymers* and *graft copolymers*. **Block copolymers** are those in which different blocks of identical monomer units alternate with each other; **graft copolymers** are those in which homopolymer branches of one monomer unit are “grafted” onto a homopolymer chain of another monomer unit.



Block copolymers are prepared by initiating the radical polymerization of one monomer as if growing homopolymer chains, followed by addition of an excess of the second monomer. Graft copolymers are made by gamma irradiation of a homopolymer chain in the presence of a second monomer. The high-energy irradiation knocks hydrogen atoms off the homopolymer chain at random points, thus generating radical sites that can initiate polymerization of the added monomer.

PROBLEM.....

- 31.10 Draw the structure of an alternating segment of butyl rubber, a copolymer of 2-methyl-1,3-butadiene and 2-methylpropene prepared under cationic conditions.

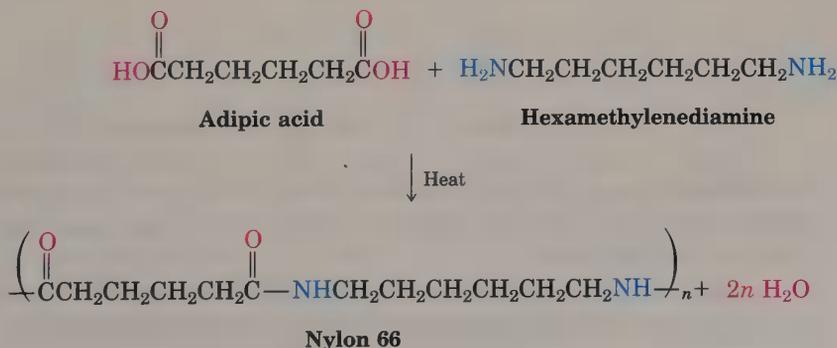
PROBLEM.....

- 31.11 One of the most important commercial applications of graft polymerization involves irradiation of polybutadiene, followed by addition of styrene. The product is used to make rubber soles for shoes. Draw the structure of a representative segment of this styrene-butadiene graft copolymer.
- .....

## 31.9 Step-Growth Polymers: Nylon

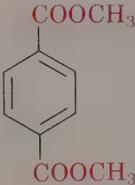
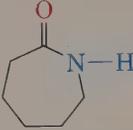
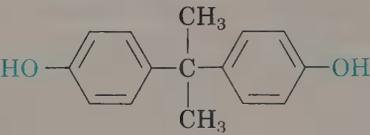
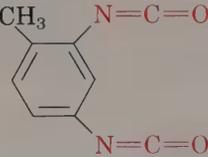
Step-growth polymers are produced by reactions between two difunctional molecules, as noted earlier. Each bond is formed in a discrete step, independent of all other bonds in the polymer, and no chain reactions are involved. The key bond-forming step is usually one of the fundamental polar reactions studied earlier, as opposed to a radical reaction. A large number of different step-growth polymers have been made, with some of the more important ones shown in Table 31.4 on page 1200.

The best-known step-growth polymers are the polyamides (**nylons**), first prepared by Wallace Carothers<sup>3</sup> at the Du Pont Company by heating diamines with diacids. For example, nylon 66 is prepared by reaction of the six-carbon adipic acid with the six-carbon hexamethylenediamine at 280°C:



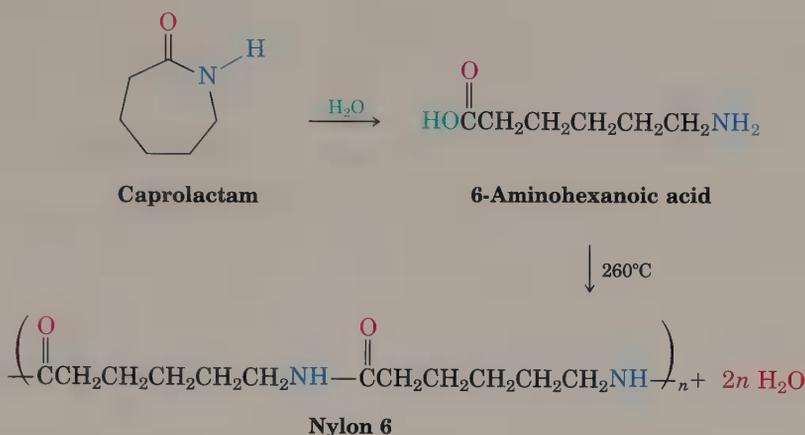
<sup>3</sup>Wallace H. Carothers (1896–1937); b. Burlington, Iowa; Ph.D. Illinois (Adams); Du Pont Company.

TABLE 31.4 Some Common Step-Growth Polymers and Their Uses

Monomer name	Formula	Trade or common name of polymer	Uses
Adipic acid Hexamethylene diamine	$\text{HOOC}(\text{CH}_2)_4\text{COOH}$ $\text{H}_2\text{N}(\text{CH}_2)_6\text{NH}_2$	Nylon 66	Fibers, clothing, tire cord, bearings
Ethylene glycol	$\text{HOCH}_2\text{CH}_2\text{OH}$		
Dimethyl terephthalate		Dacron, Terylene, Mylar	Fibers, clothing, tire cord, film
Caprolactam			
Diphenyl carbonate	$\text{C}_6\text{H}_5\text{OCOOC}_6\text{H}_5$	Lexan, polycarbonate	Molded articles, machine housings
Bisphenol A			
Poly(2-butene-1,4-diol)	$\text{HO}-(\text{CH}_2\text{CH}=\text{CHCH}_2)_n\text{OH}$	Polyurethane, Spandex	Foams, fibers, coatings
Tolylene diisocyanate			

Nylon 6, closely related in structure to nylon 66, is prepared by polymerization of caprolactam. Water is first added to hydrolyze caprolactam to 6-aminohexanoic acid, and strong heating then brings about dehydration and polymerization.

Nylons are used both in engineering applications and in making fibers. A combination of high impact strength and abrasion resistance makes nylon an excellent metal-substitute for bearings and gears. As fiber, nylon is used in a wide variety of applications, from clothing to tire cord to Perlon mountaineering ropes.

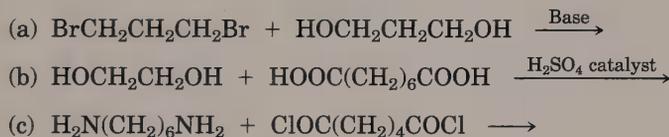


PROBLEM .....

- 31.12 Nylon is far more easily damaged by accidental spillage of acid or base than are chain-growth polymers like polyethylene. In other words, wearing nylon stockings in a chemistry laboratory can be very expensive. Explain.

PROBLEM .....

- 31.13 Draw structures of the step-growth polymers you would expect to obtain from these reactions:

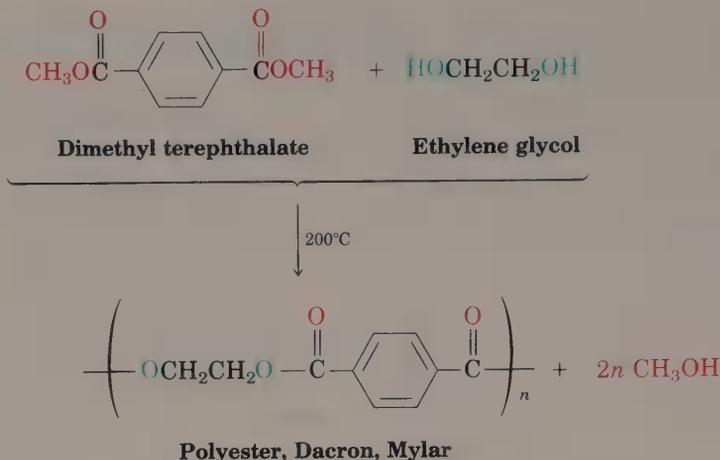


PROBLEM .....

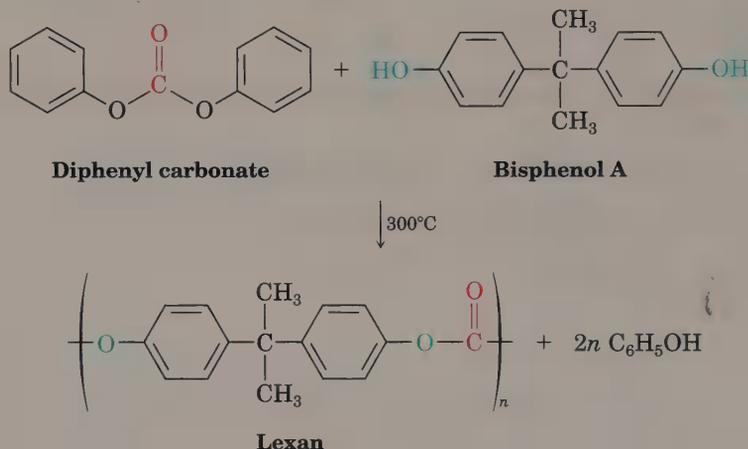
- 31.14 Kevlar, a nylon polymer prepared by reaction of 1,4-benzenedicarboxylic acid (terephthalic acid) with 1,4-diaminobenzene (*p*-phenylenediamine), is so strong that it's used to make bulletproof vests. Draw the structure of a segment of Kevlar.
- .....

## 31.10 Polyesters

Just as polyamides can be made by reaction between diacids and diamines, **polyesters** can be made by reaction between diacids and dialcohols. The most generally useful polyester is made by ester exchange reaction between dimethyl terephthalate and ethylene glycol. The product is widely used under the trade name Dacron to make clothing fiber and tire cord, and is used under the name Mylar to make plastic film. The tensile strength of poly(ethylene terephthalate) film is nearly equal to that of steel, and the film is unusually flex- and tear-resistant.



Lexan, a polycarbonate prepared from diphenyl carbonate and bisphenol A, is another commercially valuable polyester. Lexan has an unusually high impact strength, making it valuable for use in machinery housings, telephones, and bicycle safety helmets.

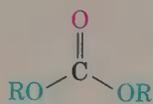


PROBLEM.....

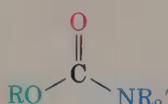
- 31.15 Draw the structure of the polymer you would expect to obtain from reaction of dimethyl terephthalate with a triol such as glycerol. What structural feature would this new polymer have that was not present in Dacron? How do you think this new feature would affect the properties of the polymer?
- .....

## 31.11 Polyurethanes

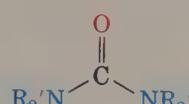
A **urethane** is a carbonyl-containing functional group in which the carbonyl carbon is bound both to an ether oxygen and to an amine nitrogen. As such, a urethane can be considered intermediate between a carbonate and a urea:



A carbonate

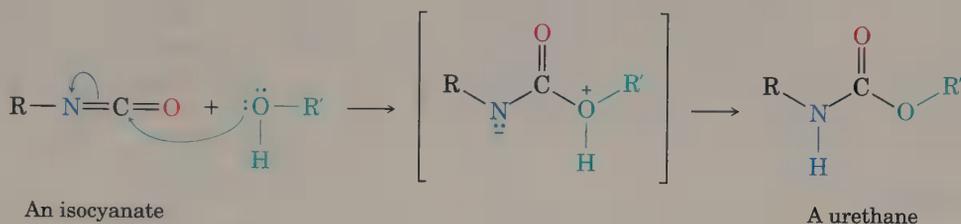


A urethane

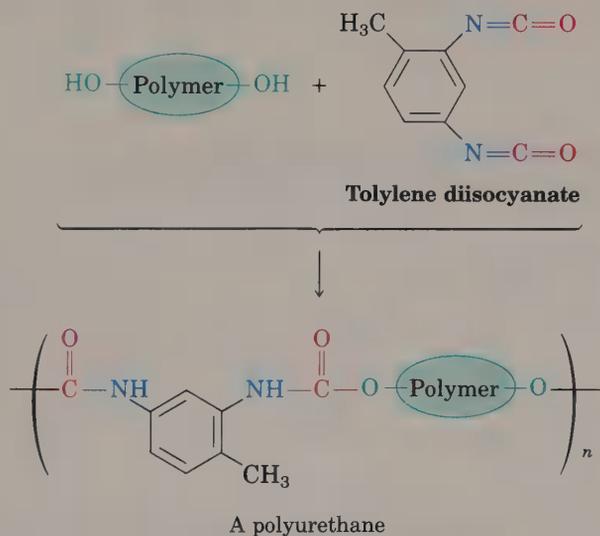


A urea

Urethanes are prepared by nucleophilic addition of alcohols to isocyanates:

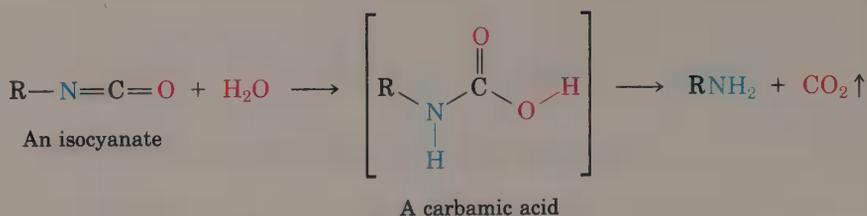


**Polyurethanes**, polymers containing urethane linkages, are prepared by reaction between diols and diisocyanates. The diol is itself a low-molecular-weight (mol wt  $\approx$  1000) polymer with hydroxyl end groups; the diisocyanate is often tolylene diisocyanate.



A number of different kinds of polyurethanes are produced, depending on the nature of the polymeric alcohol used and on the degree of cross-linking achieved. One major use of polyurethane is in the stretchable spandex and Lycra fibers used for bathing suits and leotards. These polyurethanes have a fairly low degree of cross-linking so that the resultant polymer is soft and elastic.

A second major use of polyurethanes is in foams. Foaming occurs when a small amount of water is added during polymerization. Water adds to isocyanate groups giving carbamic acids, which spontaneously lose  $\text{CO}_2$ , thus generating the foam bubbles.



Polyurethane foams generally have a higher amount of cross-linking than polyurethane fibers do, an amount that can be varied by using a poly-alcohol rather than a diol as one of the reactive components. The result is a rigid but very light foam suitable for use as thermal insulation in building construction and in portable ice chests.

PROBLEM .....

- 31.16 Show the mechanism of reaction between water and an isocyanate to yield a carbamic acid.

PROBLEM .....

- 31.17 Show the mechanism of the decarboxylation of a carbamic acid to yield an amine plus  $\text{CO}_2$ . What kind of reaction is occurring?
- .....

## 31.12 Polymer Structure and Chemistry

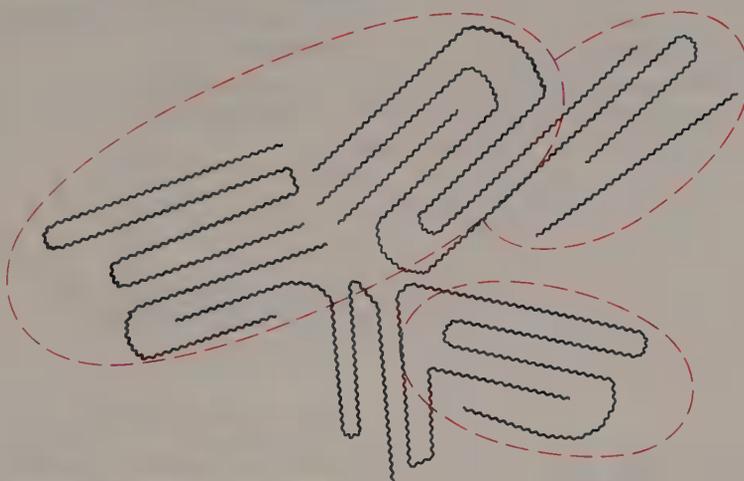
Polymers aren't really so different from other organic molecules. They're much bigger, of course, but their chemistry is the same as that of analogous small molecules. The chemistry of polymers is the familiar chemistry of functional groups; molecular size plays little role. Thus, the ester linkages of a polyester such as Dacron are hydrolyzed by base; the aromatic rings of polystyrene undergo typical electrophilic aromatic substitution reactions; and the alkane chains of polyethylene undergo radical-initiated halogenation.

The major differences between small and very large organic molecules are in structure and in physical properties. Here too, though, the bulk structures and properties of polymers are the result of the same intermolecular forces that operate in small molecules.

The most important intermolecular forces between non-cross-linked polymer chains are van der Waals forces. These forces, which are due to weak attractive interactions between transient dipoles in nearby molecules (Section 3.6), are the same as those that act between small molecules in solution or in the solid state.

Since van der Waals forces operate only at close distances, they're strongest in those polymers like linear polyethylene in which chains can line up in a regular, close-packed way. Many polymers, in fact, have regions that are essentially crystalline. These regions, called **crystallites**, consist

of highly ordered portions in which the zigzag polymer chains are bound together by van der Waals forces (Figure 31.5).



**Figure 31.5** Crystallites in linear polyethylene. The long polymer chains are arranged in parallel lines in the crystallite regions.

As you might expect, polymer crystallinity is strongly affected by the steric requirements of substituent groups on the chains. Poly(methyl methacrylate), for instance, is noncrystalline because the chains can't pack closely together in a regular way, but linear polyethylene is highly crystalline.

**PROBLEM** .....

- 31.18 What product would you expect to obtain from catalytic hydrogenation of natural rubber? Would the product be syndiotactic, atactic, or isotactic?
- .....

## 31.13 Polymer Structure and Physical Properties

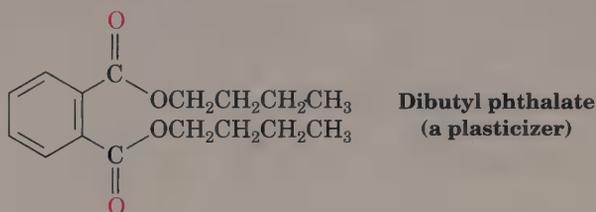
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Classification of synthetic polymers according to their physical properties is a useful exercise because it makes possible a rough correlation between structure and property. In general, we can divide polymers into four major categories: thermoplastics, fibers, elastomers, and thermosetting resins.

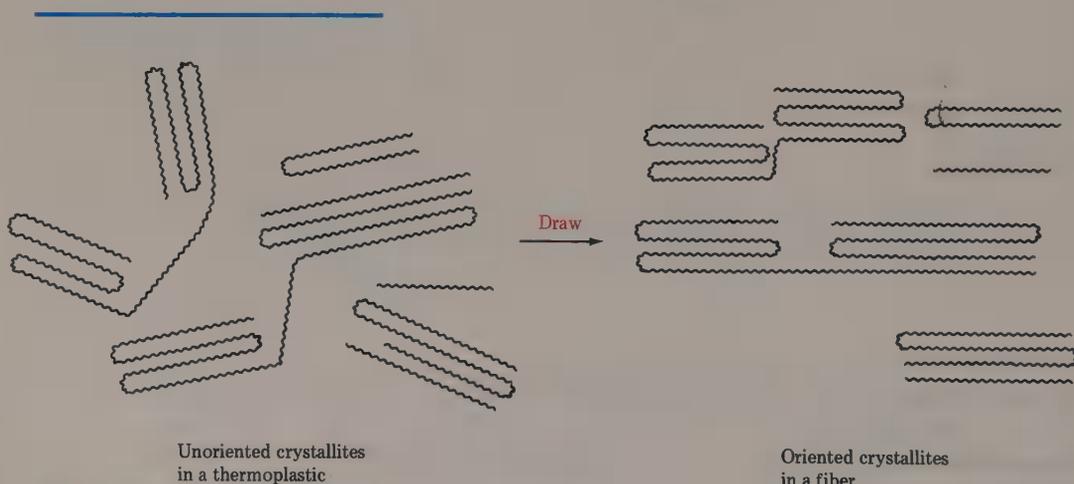
**Thermoplastics** are the polymers most people think of when the word *plastic* is mentioned. These polymers are hard at room temperature but become soft and viscous when heated. As a result, they can be molded into toys, beads, telephone housings, or into any of a thousand other items.

Because thermoplastics have little or no cross-linking, the individual chains can slip past one another on heating. Some thermoplastic polymers such as poly(methyl methacrylate) (Plexiglas) are amorphous and noncrystalline; others, such as polyethylene and nylon, are partially crystalline.

**Plasticizers**—small organic molecules that act as lubricants between chains—are usually added to plastics to keep them from becoming brittle at room temperature. Dialkyl phthalates are commonly used for this purpose and, in the past few decades, have become among the most widely dispersed of all environmental pollutants. Phthalate plasticizers have even been detected in the fat of antarctic penguins.



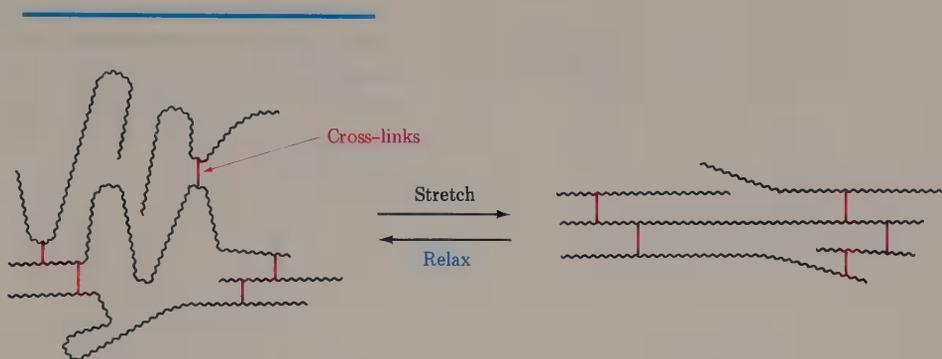
**Fibers** are thin threads produced by extruding a molten polymer through small holes in a die, or *spinneret*. The fibers are then cooled and drawn out. Drawing has the effect of orienting the crystallite regions along the axis of the fiber, a process that adds considerable tensile strength (Figure 31.6). Nylon, Dacron, and polyethylene all have the semicrystalline structure necessary for drawing into oriented fibers.



**Figure 31.6** Oriented crystallite regions in a polymer fiber.

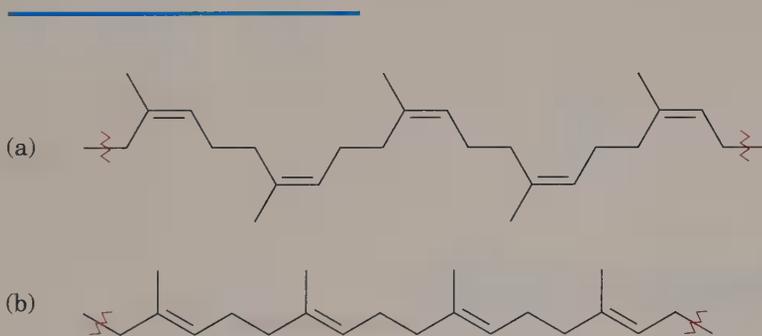
**Elastomers** are amorphous polymers that have the ability to stretch out and spring back to their original shapes. These polymers must have a modest amount of cross-linking to prevent the chains from slipping over one another, and the chains must have an irregular shape to prevent crystallite formation. When stretched, the randomly coiled chains straighten out and

orient along the direction of the pull. Van der Waals forces are too weak and too few to maintain this orientation, however, and the elastomer therefore reverts to its random-coiled state when the stretching force is released (Figure 31.7).



**Figure 31.7** Unstretched and stretched forms of an elastomer.

Natural rubber is the most obvious example of an elastomer. Rubber has the long chains and occasional cross-links needed for elasticity, but its irregular geometry prevents close packing of the chains into crystallites. Gutta-percha, by contrast, is highly crystalline and is not an elastomer (Figure 31.8).



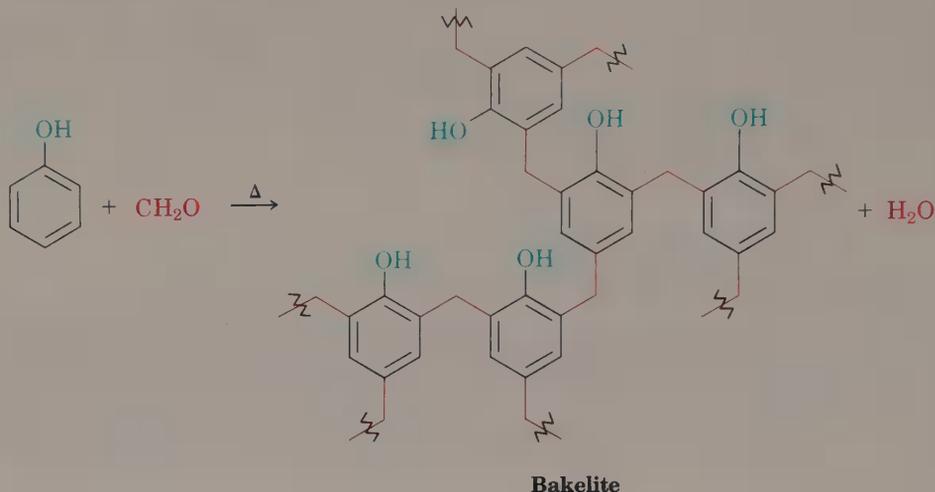
**Figure 31.8** (a) Natural rubber is elastic and noncrystalline because of its *cis* double-bond geometry, but (b) gutta-percha is nonelastic and crystalline because its geometry allows for better packing together of chains.

**Thermosetting resins** are polymers that become highly cross-linked and solidify into a hard, insoluble mass when heated. *Bakelite*, a thermosetting resin first produced in 1907 by Leo Baekeland,<sup>4</sup> has been in com-

<sup>4</sup>Leo Hendrik Baekeland (1863–1944); b. Ghent, Belgium; founder and president, Bakelite Corp., U.S. (1910–1939).

mercial use longer than any other synthetic polymer. It is widely used for molded parts, for adhesives, for coatings, and even for high-temperature applications such as missile nose cones.

Chemically, Bakelite is a *phenolic resin* produced by reaction of phenol and formaldehyde. On heating, water is eliminated, many cross-links form, and the polymer sets into a rock-like mass. The cross-linking in Bakelite and other thermosetting resins is three-dimensional and is so extensive that we can't really speak of polymer "chains." A piece of Bakelite is essentially one large molecule.



**PROBLEM** .....

- 31.19** Propose a mechanism to account for the formation of Bakelite from acid-catalyzed polymerization of phenol and formaldehyde.
- .....

## 31.14 Summary and Key Words

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Synthetic polymers can be classified as either chain-growth polymers or step-growth polymers. **Chain-growth polymers** are prepared by chain-reaction polymerization of unsaturated monomers in the presence of a radical, an anion, or a cation initiator. Radical polymerization is the most commonly used method, but alkenes such as 2-methylpropene that have electron-donating substituents on the double bond polymerize easily by a cationic route. Similarly, monomers such as methyl  $\alpha$ -cyanoacrylate that have electron-withdrawing substituents on the double bond polymerize by an anionic (Michael reaction) pathway.

Conjugated dienes such as 1,4-butadiene, isoprene, and chloroprene also undergo polymerization, giving products that have double bonds in their chains. Still another possibility is the copolymerization of two monomers to

give a product that has properties different from those of either homopolymer. **Graft copolymers** and **block copolymers** are two special kinds of polymers whose physical properties can sometimes be controlled.

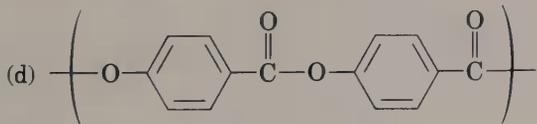
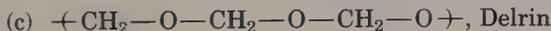
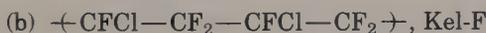
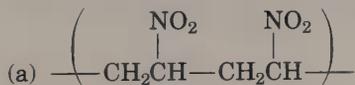
Alkene polymerization can be carried out in a much more controlled fashion if **Ziegler–Natta catalysts** are used. Ziegler–Natta polymerization minimizes the amount of chain branching in the polymer and leads to stereoregular chains—either **isotactic** (substituents on the same side of the chain) or **syndiotactic** (substituents on alternate sides of the chain), rather than **atactic** (substituents randomly disposed).

**Step-growth polymers**, the second major class of polymers, are prepared by reactions between difunctional molecules; the individual bonds in the polymer are formed independently of one another. **Polyamides** (nylons) are formed by step-growth polymerization between a diacid and a diamine; **polyesters** are formed from a diester and a diol; and **polyurethanes** are formed from a diisocyanate and a diol.

The chemistry of synthetic polymers is similar to the chemistry of small molecules with the same functional groups, but the physical properties of polymers are greatly affected by size. Polymers can be classified by physical property into four groups: **thermoplastics**, **fibers**, **elastomers**, and **thermosetting resins**. The properties of each group can be accounted for by the structure, the degree of crystallinity, the geometry, and the amount of cross-linking in the specific polymers.

## ADDITIONAL PROBLEMS .....

**31.20** Identify the monomer units in each of the following polymers.



**31.21** Identify each polymer in Problem 31.20 as either a step-growth or a chain-growth polymer.

**31.22** Cyclopentadiene undergoes thermal polymerization to yield a polymer that has no double bonds in the chain. On strong heating, this polymer breaks down to regenerate cyclopentadiene. Propose a structure for the product. (See Section 14.7.)

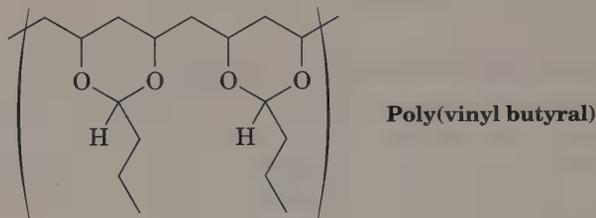
**31.23** Draw a three-dimensional representation of segments of these polymers:

- Syndiotactic polyacrylonitrile
- Atactic poly(methyl methacrylate)
- Isotactic poly(vinyl chloride)

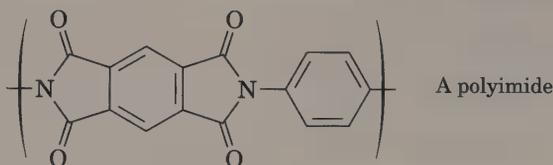
- 31.24 When styrene,  $C_6H_5CH=CH_2$ , is copolymerized in the presence of a few percent *p*-divinylbenzene, a hard, insoluble, cross-linked polymer is obtained. Show how this cross-linking of polystyrene chains occurs.
- 31.25 One method for preparing the 1,6-hexanediamine needed in nylon production starts with 1,3-butadiene. How would you accomplish this synthesis?



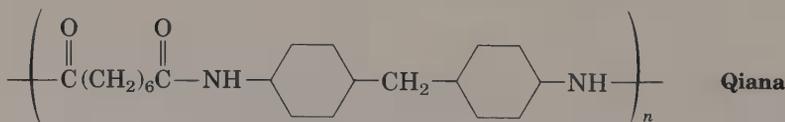
- 31.26 Nitroethylene,  $H_2C=CHNO_2$ , is a sensitive compound that must be prepared with great care. Attempted purification of nitroethylene by distillation often results in low recovery of product and a white coating on the inner walls of the distillation apparatus. Explain.
- 31.27 Poly(vinyl butyral) is used as the plastic laminate in the preparation of automobile windshield safety glass. How would you synthesize this polymer?



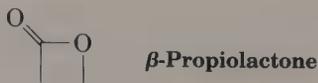
- 31.28 Polyimides having the structure shown are used as coatings on glass and plastics to improve scratch resistance. How would you synthesize a polyimide?



- 31.29 *Qiana*, a polyamide fiber with a silk-like feel, has the structure indicated. What are the monomer units used in the synthesis of *Qiana*?

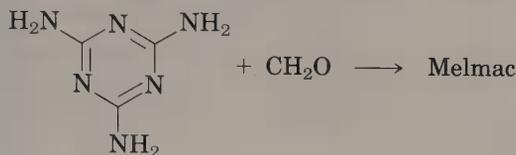


- 31.30 Starting from aniline and formaldehyde, how would you synthesize the amine necessary for the preparation of *Qiana* (Problem 31.29)?
- 31.31 What is the structure of the polymer produced by treatment of  $\beta$ -propiolactone with a small amount of hydroxide ion?



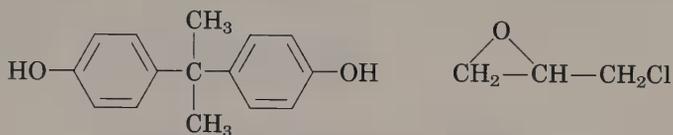
- 31.32 *Glyptal* is a highly cross-linked thermosetting resin produced by heating glycerol and phthalic anhydride (1,2-benzenedicarboxylic acid anhydride). Show the structure of a representative segment of glyptal.

- 31.33 *Melmac*, a thermosetting resin often used to make plastic dishes, is prepared by heating melamine with formaldehyde. Review the structure of Bakelite in Section 31.13, and then propose a structure for Melmac.



Melamine

- 31.34 Epoxy adhesives are cross-linked resins prepared in two steps. The first step involves  $\text{S}_{\text{N}}2$  reaction of the disodium salt of bisphenol A with epichlorohydrin to form a low-molecular-weight prepolymer. This prepolymer is then "cured" into a cross-linked resin by treatment with a triamine such as  $\text{H}_2\text{NCH}_2\text{CH}_2\text{NHCH}_2\text{CH}_2\text{NH}_2$ .

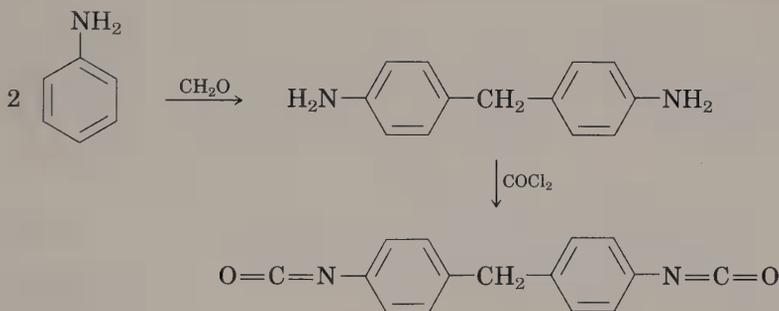


Bisphenol A

Epichlorohydrin

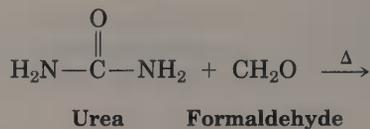
- (a) What is the structure of the prepolymer?  
 (b) How does addition of the triamine result in cross-linking?

- 31.35 The polyurethane foam used for home insulation uses methanediisocyanate (MDI) as monomer. The MDI is prepared by reaction of aniline with formaldehyde, followed by treatment with phosgene,  $\text{COCl}_2$ . Propose mechanisms for the two steps.

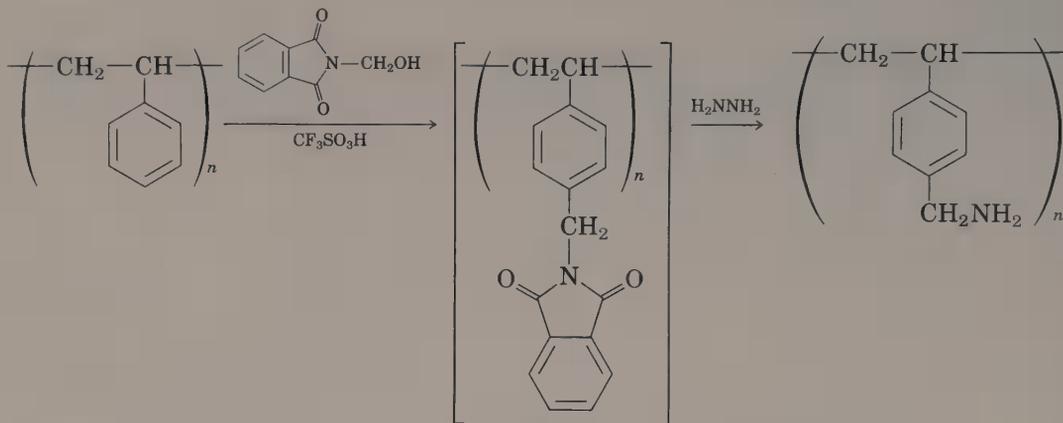


MDI

- 31.36 Write the structure of a representative segment of polyurethane prepared by reaction of ethylene glycol with MDI (Problem 31.35).
- 31.37 Urea-formaldehyde resins have been prepared commercially for over 60 years for a variety of purposes, including use as adhesives and insulating foams. For example, the smoking salons of the great hydrogen-filled dirigibles of the 1930s were insulated with urea-formaldehyde foams, and urea-formaldehyde adhesives were used extensively in boats and aircraft during World War II. The structure of the urea-formaldehyde polymer is highly cross-linked, like that of Bakelite (Section 31.13). Propose a structure for this polymer.



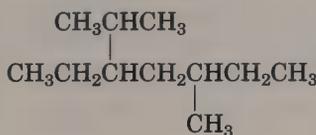
- 31.38 An improved polymeric resin used for Merrifield solid-phase peptide synthesis (Section 27.12) is prepared by treating polystyrene with *N*-(hydroxymethyl)phthalimide and trifluoromethanesulfonic acid, followed by reaction with hydrazine. Show how these steps occur.



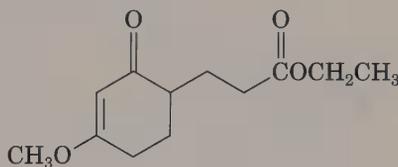
# Appendix A

## Nomenclature of Polyfunctional Organic Compounds

Judging from the number of incorrect names that appear in the chemical literature, it's probably safe to say that relatively few practicing organic chemists are fully conversant with the rules of organic nomenclature. Simple hydrocarbons and monofunctional compounds present few difficulties because the basic rules for naming such compounds are logical and easy to understand. Problems, however, are often encountered with polyfunctional compounds. Whereas most chemists could correctly identify hydrocarbon **1** as 3-ethyl-2,5-dimethylheptane, rather few could correctly identify polyfunctional compound **2**. Should we consider **2** as an ether? As an ethyl ester? As a ketone? As an alkene? It is, of course, all four, but it has only one correct name: ethyl 3-(4-methoxy-2-oxo-3-cyclohexenyl)propanoate.



1. 3-Ethyl-2,5-dimethylheptane



2. Ethyl 3-(4-methoxy-2-oxo-3-cyclohexenyl)propanoate

Naming polyfunctional organic compounds isn't really much harder than naming monofunctional ones. All that's required is a knowledge of nomenclature for monofunctional compounds and a set of additional rules. In the following discussion, it's assumed that you have a good command of the rules of nomenclature for monofunctional compounds that were given throughout the text as each new functional group was introduced. A list of where these rules can be found is shown in Table A.1.

TABLE A.1 Where to Find Nomenclature Rules for Simple Functional Groups

<i>Functional group</i>	<i>Text section</i>	<i>Functional group</i>	<i>Text section</i>
Acid anhydrides	21.1	Amines	25.1
Acid halides	21.1	Aromatic compounds	15.2
Alcohols	17.1	Carboxylic acids	20.1
Aldehydes	19.2	Cycloalkanes	3.8
Alkanes	3.4	Esters	21.1
Alkenes	6.3	Ethers	18.1
Alkyl halides	10.1	Ketones	19.2
Alkynes	8.2	Nitriles	21.1
Amides	21.1	Phenols	26.4

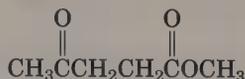
The name of a polyfunctional organic molecule has four parts:

1. **Suffix**—the part that identifies the principal functional-group class to which the molecule belongs.
2. **Parent**—the part that identifies the size of the main chain or ring.
3. **Substituent prefixes**—parts that identify what substituents are located on the main chain or ring.
4. **Locants**—numbers that tell where substituents are located on the main chain or ring.

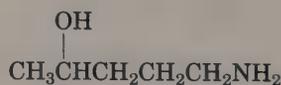
To arrive at the correct name for a complex molecule, you must identify the four name parts and then express them in the proper order and format. Let's look at the four parts.

## The Suffix—Functional-Group Precedence

A polyfunctional organic molecule can contain many different kinds of functional groups, but for nomenclature purposes, we must choose just one suffix. It's not correct to use two suffixes. Thus, keto ester **3** must be named either as a ketone with an *-one* suffix or as an ester with an *-oate* suffix but can't be named as an *-onoate*. Similarly, amino alcohol **4** must be named either as an alcohol (*-ol*) or as an amine (*-amine*) but can't properly be named as an *-olamine*. The only exception to this rule is in naming compounds that have double or triple bonds. For example, the unsaturated acid  $\text{H}_2\text{C}=\text{CHCH}_2\text{COOH}$  is 3-butenic acid, and the acetylenic alcohol  $\text{HC}\equiv\text{CCH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$  is 5-hexyn-1-ol.



3. Named as an ester with a keto (oxo) substituent: **methyl 4-oxopentanoate**



4. Named as an alcohol with an amino substituent: **5-amino-2-pentanol**

How do we choose which suffix to use? Functional groups are divided into two classes, **principal groups** and **subordinate groups**, as shown in Table A.2. Principal groups are those that may be cited either as prefixes or as suffixes, whereas

subordinate groups are those that may be cited only as prefixes. Within the principal groups, an order of precedence has been established. The proper suffix for a given compound is determined by identifying all of the functional groups present and then choosing the principal group of highest priority. For example, Table A.2 indicates that keto ester **3** must be named as an ester rather than as a ketone, since an ester functional group is higher in priority than a ketone is. Similarly, amino alcohol **4**

**TABLE A.2** Classification of Functional Groups for Purposes of Nomenclature<sup>a</sup>

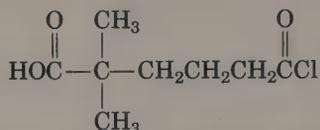
<i>Functional group</i>	<i>Name as suffix</i>	<i>Name as prefix</i>
<i>Principal groups</i>		
Carboxylic acids	-oic acid -carboxylic acid	carboxy
Acid anhydrides	-oic anhydride -carboxylic anhydride	
Esters	-oate -carboxylate	alkoxycarbonyl
Acid halides	-oyl halide -carbonyl halide	halocarbonyl
Amides	-amide -carboxamide	amido
Nitriles	-nitrile -carbonitrile	cyano
Aldehydes	-al -carbaldehyde	oxo
Ketones	-one	oxo
Alcohols	-ol	hydroxy
Phenols	-ol	hydroxy
Thiols	-thiol	mercapto
Amines	-amine	amino
Imines	-imine	imino
Alkenes	-ene	alkenyl
Alkynes	-yne	alkynyl
Alkanes	-ane	alkyl
<i>Subordinate groups</i>		
Ethers		alkoxy
Sulfides		alkylthio
Halides		halo
Nitro		nitro
Azides		azido
Diazo		diazo

<sup>a</sup>Principal functional groups are listed in order of decreasing priority; subordinate functional groups have no established priority order.

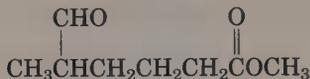
must be named as an alcohol rather than as an amine. The correct name of **3** is methyl 4-oxopentanoate, and the correct name of **4** is 5-amino-2-pentanol. Further examples are shown below.



5. Named as a cyclohexanecarboxylic acid with an oxo substituent: **4-oxocyclohexanecarboxylic acid**



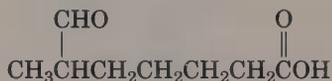
6. Named as a carboxylic acid with a chlorocarbonyl substituent: **5-chlorocarbonyl-2,2-dimethylpentanoic acid**



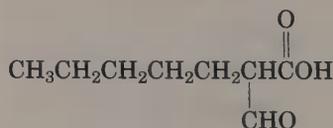
7. Named as an ester with an oxo substituent: **methyl 5-methyl-6-oxohexanoate**

## The Parent—Selecting the Main Chain or Ring

The parent or base name of a polyfunctional organic compound is usually easy to identify. If the group of highest priority is *part* of an open chain, we simply select the longest chain that contains the largest number of principal functional groups. If the highest priority group is *attached* to a ring, we use the name of that ring system as the parent. For example, compounds **8** and **9** are isomeric aldehyde acids, and both must be named as acids rather than as aldehydes according to Table A.2. The longest chain in compound **8** has seven carbons, and the substance is therefore named 6-methyl-7-oxoheptanoic acid. Compound **9** also has a chain of seven carbons, but the longest chain that contains both of the principal functional groups has only three carbons. The correct name of **9** is 3-oxo-2-pentylpropanoic acid.



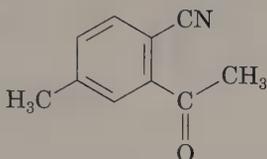
8. Named as a substituted heptanoic acid: **6-methyl-7-oxoheptanoic acid**



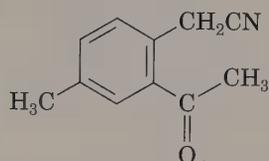
9. Named as a substituted propanoic acid: **3-oxo-2-pentylpropanoic acid**

Similar rules apply for compounds **10–13**, which contain rings. Compounds **10** and **11** are isomeric keto nitriles, and both must be named as nitriles according to Table A.2. Substance **10** is named as a benzonitrile since the  $-\text{CN}$  functional group

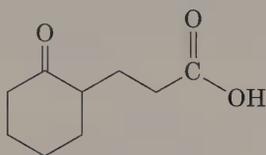
is a substituent on the aromatic ring, but substance **11** is named as an acetonitrile since the  $-\text{CN}$  functional group is on an open chain. The correct names are 2-acetyl-4-methylbenzonitrile (**10**) and (2-acetyl-4-methylphenyl)acetonitrile (**11**). Compounds **12** and **13** are both keto acids and must be named as acids. The correct names are 3-(2-oxocyclohexyl)propanoic acid (**12**) and 2-(3-oxopropyl)cyclohexanecarboxylic acid (**13**).



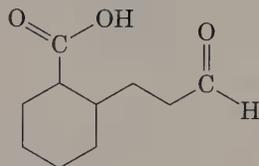
**10.** Named as a substituted benzonitrile:  
**2-acetyl-4-methylbenzonitrile**



**11.** Named as a substituted acetonitrile:  
**(2-acetyl-4-methylphenyl)acetonitrile**



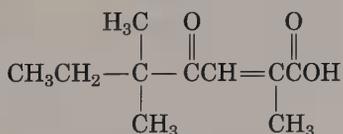
**12.** Named as a carboxylic acid:  
**3-(2-oxocyclohexyl)propanoic acid**



**13.** Named as a carboxylic acid:  
**2-(3-oxopropyl)cyclohexanecarboxylic acid**

## The Prefixes and Locants

With the suffix and parent name established, the next step is to identify and number all substituents on the parent chain or ring. These substituents include all alkyl groups and all functional groups other than the one cited in the suffix. For example, compound **14** contains three different functional groups (carboxyl, keto, and double bond). Because the carboxyl group is highest in priority, and because the longest chain containing the functional groups is seven carbons long, **14** is a heptenoic acid. In addition, the main chain has an oxo (keto) substituent and three methyl groups. Numbering from the end nearer the highest-priority functional group, we find that **14** is 2,5,5-trimethyl-4-oxo-2-heptenoic acid. Note that the final *-e* of heptene is deleted in the word *heptenoic*. This deletion occurs only when the name would have two adjacent vowels (thus, *heptenoic* has the final *e* deleted, but *heptenenitrile* retains the *-e*). Look back at some of the other compounds we've named to see other examples of how prefixes and locants are assigned.

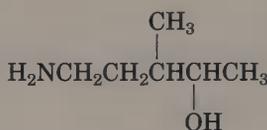


**14.** Named as a heptenoic acid:  
**2,5,5-trimethyl-4-oxo-2-heptenoic acid**

## Writing the Name

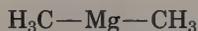
Once the name parts have been established, the entire name is written out. Several additional rules apply:

1. *Order of prefixes* When the substituents have been identified, the main chain has been numbered, and the proper multipliers such as *di-* and *tri-* have been assigned, the name is written with the substituents listed in alphabetical, rather than numerical, order. Multipliers such as *di-* and *tri-* are not used for alphabetization purposes, but the prefix *iso-* is used.

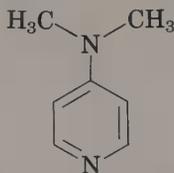


15. **5-Amino-3-methyl-2-pentanol**  
(not 3-methyl-5-amino-2-pentanol)

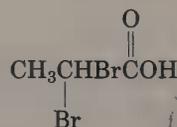
2. *Use of hyphens; single- and multiple-word names* The general rule in such cases is to determine whether the principal functional group is itself an element or compound. If it is, then the name is written as a single word; if it isn't, then the name is written as multiple words. For example, methylbenzene (one word) is correct because the parent—benzene—is itself a compound. Diethyl ether, however, is written as two words because the parent—ether—is a class name rather than a compound name. Some further examples are shown:



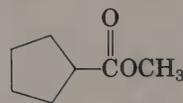
16. **Dimethylmagnesium**  
(one word, since magnesium is an element)



18. **4-(Dimethylamino)pyridine**  
(one word, since pyridine is a compound)



17. **2-Bromopropanoic acid**  
(two words, since "acid" is not a compound)

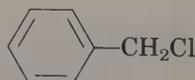


19. **Methyl cyclopentanecarboxylate**

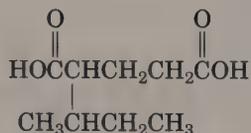
3. *Parentheses* Parentheses are used to denote complex substituents when ambiguity would otherwise arise. For example, chloromethylbenzene has two substituents on a benzene ring, but (chloromethyl)benzene has only one complex substituent. Note that the expression in parentheses is not set off by hyphens from the rest of the name.



20. **p-Chloromethylbenzene**



21. **(Chloromethyl)benzene**

**22. 2-(1-Methylpropyl)pentanedioic acid**

(The 1-methylpropyl group is a complex substituent on C2 of the main chain.)

## Additional Reading

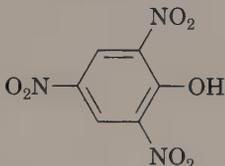
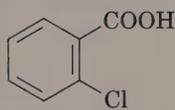
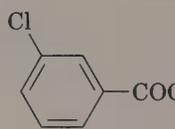
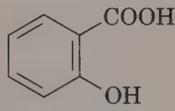
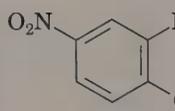
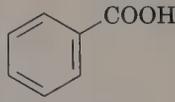
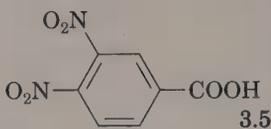
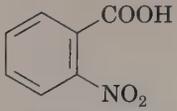
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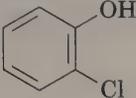
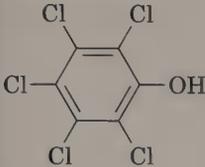
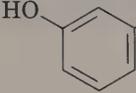
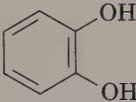
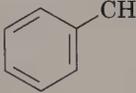
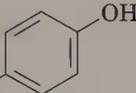
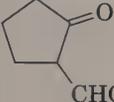
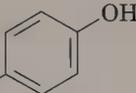
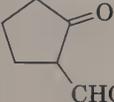
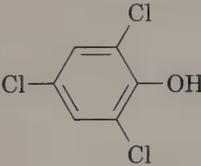
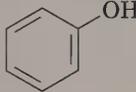
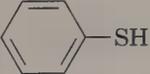
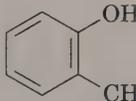
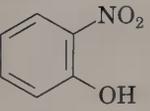
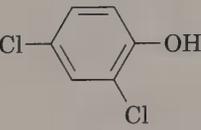
Further explanations of the rules of organic nomenclature can be found in the following references:

1. J. H. Fletcher, O. C. Dermer, and R. B. Fox, "Nomenclature of Organic Compounds: Principles and Practice," Advances in Chemistry Series No. 126, American Chemical Society, Washington, DC, 1974.
2. "Nomenclature of Organic Chemistry, Sections A, B, C, D, E, F, and H," International Union of Pure and Applied Chemistry, Pergamon Press, Oxford, 1979.
3. J. G. Traynham, "Organic Nomenclature: A Programmed Introduction," Prentice-Hall, Englewood Cliffs, NJ, 1985.

# Appendix B

## Acidity Constants for Some Organic Compounds

Compound	pK <sub>a</sub>	Compound	pK <sub>a</sub>	Compound	pK <sub>a</sub>
CH <sub>3</sub> SO <sub>3</sub> H	-1.8	CH <sub>2</sub> F <sub>2</sub> COOH	2.7	CH <sub>3</sub> OCH <sub>2</sub> COOH	3.6
CH(NO <sub>2</sub> ) <sub>3</sub>	0.1	CH <sub>2</sub> ClCOOH	2.8	CH <sub>3</sub> COCH <sub>2</sub> COOH	3.6; 7.8
	0.3	HOOCCH <sub>2</sub> COOH	2.8; 5.6	HOCH <sub>2</sub> COOH	3.7
CCl <sub>3</sub> COOH	0.5	CH <sub>2</sub> BrCOOH	2.9	HCOOH	3.7
CF <sub>3</sub> COOH	0.5		3.0		3.8
CBr <sub>3</sub> COOH	0.7		3.0		4.0
HOCC≡CCOOH	1.2; 2.5	CH <sub>2</sub> ICOOH	3.2	CH <sub>2</sub> BrCH <sub>2</sub> COOH	4.0
HOCCOOH	1.2; 3.7	CHOCOAH	3.2		4.1
CHCl <sub>2</sub> COOH	1.3		3.4		4.2
CH <sub>2</sub> (NO <sub>2</sub> )COOH	1.3		3.5	H <sub>2</sub> C=CHCOOH	4.2
HC≡CCOOH	1.9	HSCH <sub>2</sub> COOH	3.5; 10.2		
Z HOOCCH=CHCOOH	1.9; 6.3	CH <sub>2</sub> (NO <sub>2</sub> ) <sub>2</sub>	3.6		
	2.4				
CH <sub>3</sub> COCOAH	2.4				
NCCH <sub>2</sub> COOH	2.5				
CH <sub>3</sub> C≡CCOOH	2.6				

Compound	pK <sub>a</sub>	Compound	pK <sub>a</sub>	Compound	pK <sub>a</sub>
HOOCH <sub>2</sub> CH <sub>2</sub> COOH	4.2; 5.7	CH <sub>3</sub> CO <sub>3</sub> H	8.2	CH <sub>3</sub> CHO	13.5
HOOCH <sub>2</sub> CH <sub>2</sub> CH <sub>2</sub> COOH	4.3; 5.4		8.5	(CH <sub>3</sub> ) <sub>2</sub> CHCHO	13.8
	4.5	CH <sub>3</sub> CH <sub>2</sub> NO <sub>2</sub>	8.5	HOCH <sub>2</sub> CH(OH)CH <sub>2</sub> OH	14.1
CH <sub>3</sub> COOH	4.7		8.7	CH <sub>2</sub> ClCH <sub>2</sub> OH	14.3
H <sub>2</sub> C=C(CH <sub>3</sub> )COOH	4.7	CH <sub>3</sub> COCH <sub>2</sub> COCH <sub>3</sub>	9.0		15.0
CH <sub>3</sub> CH <sub>2</sub> COOH	4.8		9.3; 11.1	CH <sub>3</sub> OH	15.1
(CH <sub>3</sub> ) <sub>3</sub> CCOOH	5.0		9.3; 12.6		15.4
CH <sub>3</sub> COCH <sub>2</sub> NO <sub>2</sub>	5.1		5.3	H <sub>2</sub> C=CHCH <sub>2</sub> OH	15.5
	5.3		9.4	CH <sub>3</sub> CH <sub>2</sub> OH	15.9
O <sub>2</sub> NCH <sub>2</sub> COOCH <sub>3</sub>	5.8		9.9; 11.5	CH <sub>3</sub> CH <sub>2</sub> CH <sub>2</sub> OH	16.1
	5.8		9.9; 11.5	CH <sub>3</sub> COCH <sub>2</sub> Br	16.1
	5.8	CH <sub>3</sub> COCH <sub>2</sub> SOCH <sub>3</sub>	10.0		16.7
	6.2		10.0	(CH <sub>3</sub> ) <sub>2</sub> CHOH	17.1
	6.6		10.3	(CH <sub>3</sub> ) <sub>3</sub> COH	19.2
HCO <sub>3</sub> H	7.1	CH <sub>3</sub> NO <sub>2</sub>	10.3	CH <sub>3</sub> COCH <sub>3</sub>	20.0
	7.2	CH <sub>3</sub> SH	10.3		23
(CH <sub>3</sub> ) <sub>2</sub> CHNO <sub>2</sub>	7.7	CH <sub>3</sub> COCH <sub>2</sub> COOCH <sub>3</sub>	10.6	CH <sub>3</sub> COOCH <sub>2</sub> CH <sub>3</sub>	25
	7.8	CH <sub>3</sub> COCHO	11.0	HC≡CH	25
		CH <sub>2</sub> (CN) <sub>2</sub>	11.2	CH <sub>3</sub> CN	25
		CCl <sub>3</sub> CH <sub>2</sub> OH	12.2	CH <sub>3</sub> SO <sub>2</sub> CH <sub>3</sub>	28
		Glucose	12.3	(C <sub>6</sub> H <sub>5</sub> ) <sub>3</sub> CH	32
		(CH <sub>3</sub> ) <sub>2</sub> C=NOH	12.4	(C <sub>6</sub> H <sub>5</sub> ) <sub>2</sub> CH <sub>2</sub>	34
		CH <sub>2</sub> (COOCH <sub>3</sub> ) <sub>2</sub>	12.9	CH <sub>3</sub> SOCH <sub>3</sub>	35
		CHCl <sub>2</sub> CH <sub>2</sub> OH	12.9	NH <sub>3</sub>	35
		CH <sub>2</sub> (OH) <sub>2</sub>	13.3	CH <sub>3</sub> CH <sub>2</sub> NH <sub>2</sub>	35
				(CH <sub>3</sub> CH <sub>2</sub> ) <sub>2</sub> NH	36
					41
					43
				H <sub>2</sub> C=CH <sub>2</sub>	44
				CH <sub>4</sub>	~60

An acidity list covering more than 5000 organic compounds has been published: E. P. Serjeant and B. Dempsey (eds.), "Ionization Constants of Organic Acids in Aqueous Solution," IUPAC Chemical Data Series No. 23, Pergamon Press, Oxford, 1979.

# Appendix C

## Glossary

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**Absolute configuration** (Section 9.6): The actual three-dimensional structure of a chiral molecule. Absolute configurations are specified verbally by the Cahn–Ingold–Prelog *R,S* convention and are represented on paper by Fischer projections.

**Absorption spectrum** (Section 12.8): A plot of wavelength of incident light versus amount of light absorbed. Organic molecules show absorption spectra in both the infrared and the ultraviolet regions of the electromagnetic spectrum. By interpreting these spectra, we can obtain useful structural information about the sample. (See Infrared spectrum, Ultraviolet spectroscopy.)

**Acetal** (Section 19.14): A functional group consisting of two ether-type oxygen atoms bound to the same carbon,  $R_2C(OR')_2$ . Acetals are often used as protecting groups for ketones and aldehydes since they are stable to basic and nucleophilic reagents but can be easily removed by acidic hydrolysis.

**Achiral** (Section 9.2): Having a lack of handedness. A molecule is achiral if it has a plane of symmetry and is thus superimposable on its mirror image. (See Chiral.)

**Activating group** (Section 16.5): An electron-donating group such as hydroxyl ( $-OH$ ) or amino ( $-NH_2$ ) that increases the reactivity of an aromatic ring toward electrophilic aromatic substitution. All activating groups are ortho- and para-directing.

**Activation energy,  $\Delta G^\ddagger$**  (Section 5.9): The difference in energy levels between ground state and transition state. The amount of activation energy required by a reaction determines the rate at which the reaction proceeds. The majority of organic reactions have activation energies of 10–25 kcal/mol.

**Acylation** (Section 16.4): The introduction of an acyl group,  $-COR$ , onto a molecule. For example, acylation of an alcohol yields an ester ( $R'OH \rightarrow R'OCOR$ ), acylation of an amine yields an amide ( $R'NH_2 \rightarrow R'NHCOR$ ), and acylation of an aromatic ring yields an alkyl aryl ketone ( $ArH \rightarrow ArCOR$ ).

**Acylium ion** (Section 16.4): A resonance-stabilized carbocation in which the positive charge is located at a carbonyl-group carbon,  $R-C^+=O \longleftrightarrow R-C\equiv O^+$ . Acylium ions are strongly electrophilic and are involved as intermediates in Friedel–Crafts acylation reactions.

**1,4 Addition** (Sections 14.5 and 19.18): Addition of a reagent to the ends of a conjugated pi system. Conjugated dienes yield 1,4 adducts when treated with electrophiles such as HCl. Conjugated enones yield 1,4 adducts when treated with nucleophiles such as cyanide ion.

**Adrenocortical hormone** (Section 28.8): A steroid hormone secreted by the adrenal glands. There are two types of adrenocortical hormones: mineralocorticoids and glucocorticoids.

**Aldaric acid** (Section 24.8): The dicarboxylic acid resulting from oxidation of an aldose.

**Alditol** (Section 24.8): The polyalcohol resulting from reduction of the carbonyl group of a sugar.

**Aldonic acid** (Section 24.8): The monocarboxylic acid resulting from mild oxidation of an aldose.

**Alicyclic** (Section 3.7): Referring to an aliphatic cyclic hydrocarbon such as a cycloalkane or cycloalkene.

**Aliphatic** (Section 3.2): Referring to a nonaromatic hydrocarbon such as a simple alkane, alkene, or alkyne.

**Alkaloid** (Section 25.10): A naturally occurring compound that contains a basic amine functional group. Morphine is an example of an alkaloid.

**Alkylation** (Sections 8.9, 16.3, 18.4, and 22.9): Introduction of an alkyl group onto a molecule. For example, certain aromatic rings can be alkylated to yield arenes ( $ArH \rightarrow ArR$ ), alkoxide anions can be alkylated to yield ethers ( $R'O^- \rightarrow R'OR$ ), and enolate anions can be alkylated to yield  $\alpha$ -substituted carbonyl compounds ( $R'_2C=C(R')O^- \rightarrow RR'_2C-COR'$ ).

**Allylic** (Section 10.5): Used to refer to the position next to a double bond. For example,  $H_2C=CHCH_2Br$  is an allylic bromide, and an allylic radical is a conjugated, resonance-stabilized species in which the unpaired electron is in a *p* orbital next to a double bond ( $C=C-C\cdot \longleftrightarrow \cdot C-C=C$ ).

**Androgen** (Section 28.8): A male steroid sex hormone.

**Angle strain** (Section 4.4): The strain introduced into a molecule when a bond angle is deformed from its ideal value. Angle strain is particularly important in small-ring cycloalkanes where it results from compression of bond angles to less than their ideal tetrahedral values. For example, cyclopropane has approximately 22 kcal/mol angle strain because of bond deformation from the 109° tetrahedral angle to 60°.

**Anomers** (Section 24.6): Cyclic stereoisomers of sugars that differ only in their configurations at the hemiacetal (anomeric) carbon.

**Antarafacial** (Section 30.8): A word used to describe the geometry of pericyclic reactions. An antarafacial reaction is one that takes place on opposite faces of the two ends of a pi electron system. (See Suprafacial.)

**Anti conformation** (Section 4.3): The geometric arrangement around a carbon–carbon single bond, in which the two largest substituents are 180° apart as viewed in a Newman projection.

**Anti stereochemistry** (Section 7.2): Referring to opposite sides of a double bond or molecule. An anti addition reaction is one in which the two ends of the double bond are attacked from different sides. For example, addition of  $\text{Br}_2$  to cyclohexene yields *trans*-1,2-dibromocyclohexane, the product of anti addition. An anti elimination reaction is one in which the two groups leave from opposite sides of the molecule. (See Syn stereochemistry.)

**Antibonding orbital** (Section 1.7): A molecular orbital that is higher in energy than the atomic orbitals from which it is formed.

**Anticodon** (Section 29.15): A sequence of three bases on tRNA that read the codons on mRNA and bring the correct amino acids into position for protein synthesis.

**Aromaticity** (Chapter 15): The special characteristics of cyclic conjugated pi electron systems that result from their electronic structures. These characteristics include unusual stability, the presence of a ring current in the  $^1\text{H}$  NMR spectrum, and a tendency to undergo substitution reactions rather than addition reactions on treatment with electrophiles. Aromatic molecules must be planar, cyclic, conjugated species that have  $4n + 2$  pi electrons.

**Asymmetric center** (Section 9.2): See stereogenic center.

**Atactic polymers** (Section 31.6): Chain-growth polymers that have a random stereochemical arrangement of substituents on the polymer backbone. These polymers result from high-temperature radical-initiated polymerization of alkene monomers.

**Aufbau principle** (Section 1.3): A guide for determining the ground-state electronic configuration of elements by filling the lowest energy orbitals first.

**Axial bond** (Section 4.10): A bond to chair cyclohexane that lies along the ring axis perpendicular to the rough plane of the ring. (See Equatorial bond.)



Axial bonds

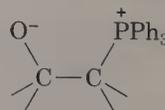
**Base peak** (Section 12.5): The most intense peak in a mass spectrum.

**Bent bonds** (Section 4.7): The bonds in small rings such as cyclopropane that bend away from the internuclear line and overlap at a slight angle, rather than head-on. Bent bonds are highly strained and highly reactive.

**Benzylic** (Sections 11.9 and 16.11): Referring to the position next to an aromatic ring. For example, a benzylic cation is a resonance-stabilized, conjugated carbocation having its positive charge located on a carbon atom next to the benzene ring in a pi orbital that overlaps the aromatic pi system.

**Benzyne** (Section 16.10): An unstable intermediate having a triple bond in a benzene ring. Benzyne is implicated as intermediates in certain nucleophilic aromatic substitution reactions of aryl halides with strong bases.

**Betaine** (Section 19.16): A neutral dipolar molecule that has nonadjacent positive and negative charges. For example, the initial adducts of Wittig reagents with carbonyl compounds are betaines.

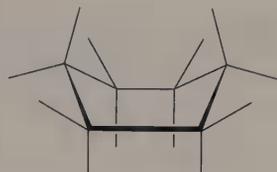


A Wittig betaine

**Bimolecular reaction** (Section 11.4): A reaction that occurs between two reagents.

**Block copolymer** (Section 31.8): A polymer consisting of alternating homopolymer blocks. Block copolymers are usually prepared by initiating chain-growth polymerization of one monomer, followed by addition of an excess of a second monomer.

**Boat cyclohexane** (Section 4.14): A three-dimensional conformation of cyclohexane that bears a slight resemblance to a boat. Boat cyclohexane has no angle strain but has a large number of eclipsing interactions that make it less stable than chair cyclohexane:



Boat cyclohexane

**Bond angle** (Section 1.8): The angle formed between two adjacent bonds.

**Bond dissociation energy** (Section 5.8): The amount of energy needed to break a bond homolytically to produce two radical fragments.

**Bond length** (Section 1.7): The equilibrium distance between the nuclei of two atoms that are bonded to each other.

**Bond strength** (Section 1.7): See Bond-dissociation energy.

**Bonding orbital** (Section 1.7): A molecular orbital that is lower in energy than the atomic orbitals from which it is formed.

**Bromohydrin** (Section 7.3): The 1,2-disubstituted bromoalcohol that is obtained by addition of HOBr to an alkene.

**Brønsted acid** (Section 2.6): A substance that donates a hydrogen ion (proton) to a base.

**Carbanion** (Section 10.9): A carbon-anion, or substance that contains a trivalent, negatively charged carbon atom ( $R_3C^-$ ). Carbanions are  $sp^3$  hybridized and have eight electrons in the outer shell of the negatively charged carbon.

**Carbene** (Section 7.11): A neutral substance that contains a divalent carbon atom having only six electrons in its outer shell ( $R_2C^:$ ).

**Carbinolamine** (Section 19.12): A molecule that contains the  $R_2C(OH)NH_2$  functional group. Carbinolamines are produced as unstable intermediates during the nucleophilic addition of amines to carbonyl groups.

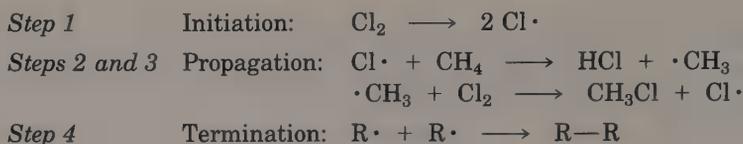
**Carbocation** (Section 6.9): A carbon cation, or substance that contains a trivalent, positively charged carbon atom having six electrons in its outer shell ( $R_3C^+$ ). Carbocations are planar and  $sp^2$  hybridized.

**Carbocycle** (Section 15.9): A cyclic molecule that has only carbon atoms in the ring. (See Heterocycle.)

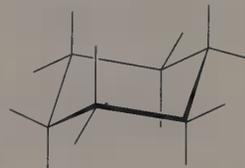
**Carbohydrate** (Chapter 24): A polyhydroxy aldehyde or polyhydroxy ketone. The name derives from the fact that glucose, the most abundant carbohydrate, has the formula  $C_6H_{12}O_6$  and was originally thought to be a "hydrate of carbon." Carbohydrates can be either simple sugars such as glucose or complex sugars such as cellulose. Simple sugars are those that can't be hydrolyzed to yield smaller molecules, whereas complex sugars are those that can be hydrolyzed to yield simpler sugars.

**Chain-growth polymer** (Section 31.1): A polymer produced by a chain-reaction procedure in which an initiator adds to a carbon-carbon double bond to yield a reactive intermediate. The chain is then built as more monomers add successively to the reactive end of the growing chain.

**Chain reaction** (Section 5.4): A reaction that, once initiated, sustains itself in an endlessly repeating cycle of propagation steps. The radical chlorination of alkanes is an example of a chain reaction that is initiated by irradiation with light and then continues in a series of propagation steps:



**Chair cyclohexane** (Section 4.9): A three-dimensional conformation of cyclohexane that resembles the rough shape of a chair. The chair form of cyclohexane, which has neither angle strain nor eclipsing strain, represents the lowest-energy conformation of the molecule:



Chair cyclohexane

**Chemical shift** (Section 13.3): The position on the NMR chart where a nucleus absorbs. By convention, the chemical shift of tetramethylsilane is arbitrarily set at zero, and all other absorptions usually occur downfield (to the left on the chart). Chemical shifts are expressed in delta units,  $\delta$ , where 1  $\delta$  equals 1 ppm of the spectrometer operating frequency. For example, 1  $\delta$  on a 60 megahertz instrument equals 60 hertz. The chemical shift of a given nucleus is related to the chemical environment of that nucleus in the molecule, thus allowing one to obtain structural information by interpreting the NMR spectrum.

**Chiral** (Section 9.2): Having handedness. Chiral molecules are those that do not have a plane of symmetry and are therefore not superimposable on their mirror image. A chiral molecule thus exists in two forms, one right-handed and one left-handed. The most common (though not the only) cause of chirality in a molecule is the presence of a carbon atom that is bonded to four different substituents. (See Achiral.)

**Chiral center:** See Stereogenic center.

**Chlorohydrin** (Section 7.3): The 1,2-disubstituted chloroalcohol that is obtained by addition of HOCl to an alkene.

**Chromatography** (Section 12.1): A technique for separating a mixture of compounds into pure components. Chromatography operates on a principle of differential adsorption whereby different compounds adsorb to a stationary support phase and are then carried along at different rates by a mobile phase.

**Cis-trans isomers** (Sections 3.9 and 6.5): Certain kinds of stereoisomers that differ in their stereochemistry about a double bond or on a ring. Cis-trans isomers are also called geometric isomers.

**Codon** (Section 29.15): A three-base sequence on the messenger RNA chain that encodes the genetic information necessary to cause specific amino acids to be incorporated into proteins. Codons on mRNA are read by complementary anticodons on tRNA.

**Concerted** (Section 30.1): Referring to a reaction that takes place in a single step without intermediates. For example, the Diels-Alder cycloaddition reaction is a concerted process.

**Configuration** (Section 9.6): The three-dimensional arrangement of atoms bonded to a stereogenic center relative to the stereochemistry of other stereogenic centers in the same molecule.

**Conformation** (Section 4.1): The exact three-dimensional shape of a molecule at any given instant, assuming that rotation around single bonds is frozen.

**Conformational analysis** (Section 4.13): A means of assessing the minimum-energy conformation of a substituted cycloalkane by totaling the steric interactions present in the molecule. Conformational analysis is particularly useful in assessing the relative stabilities of different conformations of substituted cyclohexane rings.

**Conjugate acid** (Section 2.6): The product that results from protonation of a Brønsted base.

**Conjugate addition** (Section 19.18): Addition of a nucleophile to the  $\beta$ -carbon atom of an  $\alpha,\beta$ -unsaturated carbonyl compound. (See 1,4 Addition.)

**Conjugate base** (Section 2.6): The anion that results from dissociation of a Brønsted acid.

**Conjugation** (Section 14.1): A series of alternating single and multiple bonds with overlapping  $p$  orbitals. For example, 1,3-butadiene is a conjugated diene, 3-buten-2-one is a conjugated enone, and benzene is a cyclic conjugated triene.

**Conrotatory** (Section 30.5): A term used to indicate the fact that  $p$  orbitals must rotate in the same direction during electrocyclic ring opening or ring closure. (See Disrotatory.)

**Constitutional isomers** (Sections 3.2 and 9.11): Isomers that have their atoms connected in a different order. For example, butane and 2-methylpropane are constitutional isomers.

**Copolymer** (Section 31.8): A polymer formed by chain-growth polymerization of a mixture of two or more different monomer units.

**Coupling constant** (Section 13.12): The magnitude (expressed in hertz) of the spin-spin splitting interaction between nuclei whose spins are coupled. Coupling constants are denoted  $J$ .

**Covalent bond** (Section 1.6): A bond formed by sharing electrons between two nuclei. (See Ionic bond.)

**Cracking** (Section 3.5): A process used in petroleum refining in which large alkanes are thermally cracked into smaller fragments.

**Cycloaddition** (Sections 14.7, 30.1, and 30.8): A pericyclic reaction in which two reactants add together in a single step to yield a cyclic product. The Diels-Alder reaction between a diene and a dienophile to give a cyclohexene is the best-known example of a cycloaddition.

**Deactivating group** (Section 16.5): An electron-withdrawing substituent that decreases the reactivity of an aromatic ring toward electrophilic aromatic substitution. Most deactivating groups, such as nitro, cyano, and carbonyl, are meta-directors, but halogen substituents are ortho- and para-directors.

**Decarboxylation** (Sections 20.9 and 22.9): A reaction that involves loss of carbon dioxide from the starting material.  $\beta$ -Keto acids decarboxylate particularly readily on heating.

**Degenerate orbitals** (Section 15.6): Two or more orbitals that have the same energy level.

**Dehydration** (Sections 7.1 and 17.8): A reaction that involves loss of water from the starting material. Most alcohols can be dehydrated to yield alkenes, but aldol condensation products ( $\beta$ -hydroxy ketones) dehydrate particularly readily.

**Dehydrohalogenation** (Section 7.1): A reaction that involves loss of HX from the starting material. Alkyl halides undergo dehydrohalogenation to yield alkenes on treatment with strong base.

**Delocalization** (Section 10.6): A spreading out of electron density over a conjugated pi electron system. For example, allylic cations and allylic anions are delocalized because their charges are spread out by resonance stabilization over the entire pi electron system.

**Denaturation** (Section 27.16): The physical changes that occur in proteins when secondary and tertiary structures are disrupted. Denaturation is usually brought about by heat treatment or by a change in pH and is accompanied by a loss of biological activity.

**Deshielding** (Section 13.2): An effect observed in NMR that causes a nucleus to absorb downfield (to the left) of tetramethylsilane standard. Deshielding is caused by a withdrawal of electron density from the nucleus and is responsible for the observed chemical shifts of vinylic and aromatic protons.

**Deuterium isotope effect** (Section 11.13): A tool for use in mechanistic investigations to establish whether or not a C–H bond is broken in the rate-limiting step of a reaction. Since carbon–deuterium bonds are stronger and less easily broken than carbon–protium bonds, one can measure the reaction rates of both protium- and deuterium-substituted substrates and see if they are the same or different. If they are different, a deuterium isotope effect is present, indicating that C–H bond breakage is rate-limiting.

**Dextrorotatory** (Section 9.3): A word used to describe an optically active substance that rotates the plane of polarization of plane-polarized light in a right-handed (clockwise) direction. The direction of rotation is not related to the absolute configuration of the molecule. (*See* Levorotatory.)

**Diastereomer** (Section 9.7): A term that indicates the relationship between non-mirror-image stereoisomers. Diastereomers are stereoisomers that have the same configuration at one or more stereogenic centers, but differ at other stereogenic centers.

**Diazotization** (Section 26.3): The conversion of a primary amine,  $\text{RNH}_2$ , into a diazonium salt,  $\text{RN}_2^+$ , by treatment with nitrous acid. Aryl diazonium salts are stable, but alkyl diazonium salts are extremely reactive and are rarely isolable.

**Dielectric constant** (Section 11.9): A measure of the ability of a solvent to act as an insulator of electric charge. Solvents that have high dielectric constants are highly polar and are particularly valuable in  $\text{S}_{\text{N}}1$  reactions because of their ability to stabilize the developing positive charge of the intermediate carbocation.

**Dienophile** (Section 14.8): A compound containing a double bond that can take part in the Diels–Alder cycloaddition reaction. The most reactive dienophiles are those that have electron-withdrawing groups such as nitro, cyano, or carbonyl on the double bond.

**Dipole moment,  $\mu$**  (Section 2.5): A measure of the polarity of a molecule. A dipole moment arises when the centers of gravity of positive and negative charges within a molecule do not coincide.

**Disrotatory** (Section 30.5): A term used to indicate the fact that *p* orbitals rotate in opposite directions during electrocyclic ring opening or ring closing. (*See* Conrotatory.)

**d,l form** (Section 9.11): A shorthand way of indicating the racemic modification of a compound. (*See* Racemic mixture.)

**DNA** (Section 29.9): Deoxyribonucleic acid, the biopolymer consisting of deoxyribonucleotide units linked together through phosphate–sugar bonds. DNA, which is found in the nucleus of cells, contains an organism's genetic information.

**Doublet** (Section 13.6): A two-line NMR absorption caused by spin–spin splitting when the spin of the nucleus under observation couples with the spin of a neighboring magnetic nucleus.

**Downfield** (Section 13.3): Used to refer to the left-hand portion of the NMR chart. (See Deshielding.)

**Eclipsed conformation** (Section 4.1): The geometric arrangement around a carbon–carbon single bond in which the bonds to substituents on one carbon are parallel to the bonds to substituents on the neighboring carbon as viewed in a Newman projection. For example, the eclipsed conformation of ethane has the C–H bonds on one carbon lined up with the C–H bonds on the neighboring carbon.



Eclipsed conformation

**Eclipsing strain** (Section 4.1): The strain energy in a molecule caused by electron repulsions between eclipsed bonds. Eclipsing strain is also called torsional strain.

**Elastomers** (Section 31.13): Amorphous polymers that have the ability to stretch out and then return to their previous shape. These polymers have irregular shapes that prevent crystallite formation and have little cross-linking between chains.

**Electrocyclic reaction** (Sections 30.1 and 30.5): A unimolecular pericyclic reaction in which a ring is formed or broken by a concerted reorganization of electrons through a cyclic transition state. For example, the cyclization of 1,3,5-hexatriene to yield 1,3-cyclohexadiene is an electrocyclic reaction.

**Electromagnetic spectrum** (Section 12.8): The range of electromagnetic energy, including infrared, ultraviolet, and visible radiation.

**Electron affinity** (Section 1.5): The measure of an atom's tendency to gain an electron and form an anion. Elements on the right side of the periodic table such as the halogens have higher electron affinities than do elements on the left side.

**Electronegativity** (Section 2.4): The ability of an atom to attract electrons and thereby polarize a bond. As a general rule, electronegativity increases in going across the periodic table from right to left and in going from bottom to top.

**Electrophile** (Section 5.5): An “electron-lover,” or substance that accepts an electron pair from a nucleophile in a polar bond-forming reaction.

**Electrophoresis** (Section 27.3): A technique used for separating charged organic molecules, particularly proteins and amino acids. The mixture to be separated is placed on a buffered gel or paper, and an electric potential is applied across the ends of the apparatus. Negatively charged molecules migrate toward the positive electrode, and positively charged molecules migrate toward the negative electrode.

**Elution** (Section 12.1): The removal of a substance from a chromatography column.

**Empirical formula** (Section 2.9): A formula that gives the relative proportions of elements in a compound in smallest whole numbers.

**Enantiomers** (Sections 9.1 and 9.5): Stereoisomers of a chiral substance that have a mirror-image relationship. Enantiomers must have opposite configurations at all stereogenic centers in the molecule.

**Endothermic** (Section 5.7): A term used to describe reactions that absorb heat and therefore have positive enthalpy changes. In reaction energy diagrams, the products of endothermic reactions have higher energy levels than the starting materials do.

**Entgegen (*E*)** (Section 6.6): A term used to describe the stereochemistry of a carbon–carbon double bond. The two groups on each carbon are first assigned priorities

according to the Cahn–Ingold–Prelog sequence rules, and the two carbons are then compared. If the high-priority groups on each carbon are on opposite sides of the double bond, the bond has *E* geometry. (See *Zusammen*.)

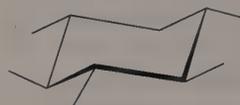
**Enthalpy change,  $\Delta H$**  (Section 5.7): The heat of reaction. The enthalpy change that occurs during a reaction is a measure of the difference in total bond energy between reactants and products.

**Entropy change,  $\Delta S$**  (Section 5.7): The change in amount of disorder. The entropy change that occurs during a reaction is a measure of the difference in disorder between reactants and products.

**Enzyme** (Chapter 27): A biological catalyst. Enzymes are large proteins that catalyze specific biochemical reactions.

**Epoxide** (Section 18.7): A three-membered-ring ether functional group.

**Equatorial bond** (Section 4.10): A bond to cyclohexane that lies along the rough equator of the ring. (See Axial bond.)



Equatorial bonds

**Equilibrium constant** (Section 2.6): A measure of the equilibrium position for a reaction. The equilibrium constant,  $K_{\text{eq}}$ , for the reaction  $A + B \rightarrow C + D$  is given by the expression

$$K_{\text{eq}} = \frac{[C][D]}{[A][B]}$$

where the numbers in brackets refer to the molar concentrations of the reactants and products.

**Essential oil** (Section 28.6): The volatile oil obtained by steam distillation of a plant extract.

**Estrogen** (Section 28.8): A female steroid sex hormone.

**Excited-state configuration** (Section 1.8): An electronic configuration having a higher energy level than the ground state has. Excited states are normally obtained by excitation of an electron from a bonding orbital to an antibonding one, such as occurs during irradiation of a molecule with light of the proper frequency.

**Exothermic** (Section 5.7): A term used to describe reactions that release heat and that therefore have negative enthalpy changes. On reaction energy diagrams, the products of exothermic reactions have energy levels lower than those of starting materials.

**Fat** (Section 28.1): A solid triacylglycerol derived from animal sources.

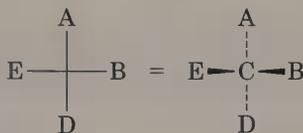
**Fibers** (Section 31.13): Thin threads produced by extruding a molten polymer through small holes in a die.

**Fibrous protein** (Section 27.13): Proteins that consist of polypeptide chains arranged side by side in long threads. These proteins are tough, insoluble in water, and are used in nature for structural materials such as hair, hooves, and fingernails.

**Fingerprint region** (Section 12.10): The complex region of the infrared spectrum from  $1500 \text{ cm}^{-1}$  to  $400 \text{ cm}^{-1}$ . If two substances have identical absorption patterns in the fingerprint region of the IR, they are almost certainly identical.

**Fischer projection** (Sections 9.13 and 24.2): A means of depicting the absolute configuration of chiral molecules on a flat page. A Fischer projection employs a cross

to represent the stereogenic center; the horizontal arms of the cross represent bonds coming out of the plane of the page, and the vertical arms of the cross represent bonds going back into the plane of the page:



Fischer projection

**Formal charge** (Section 2.3): The difference in the number of electrons owned by an atom in a molecule and by the same atom in its elemental state. The formal charge on an atom is given by the formula

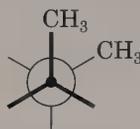
$$\text{Formal charge} = \left[ \text{Number of outer-shell electrons in free atom} \right] - \left[ \text{Number of outer-shell electrons in bonded atom} \right]$$

**Frequency** (Section 12.8): The number of electromagnetic wave cycles that travel past a fixed point in a given unit of time. Frequencies are expressed in units of cycles per second, or hertz.

**Functional group** (Section 3.1): An atom or group of atoms that is part of a larger molecule and has a characteristic chemical reactivity. Functional groups display the same chemistry in all molecules of which they are a part.

**Gated-decoupled mode** (Section 13.5): A mode of  $^{13}\text{C}$  NMR spectrometer operation in which all one-carbon resonances are of equal intensity. Operating in this mode allows one to integrate the spectrum to find out how many of each kind of carbon atom are present.

**Gauche conformation** (Section 4.3): The conformation of butane in which the two methyl groups lie  $60^\circ$  apart as viewed in a Newman projection. This conformation has 0.9 kcal/mol steric strain:



Gauche conformation

**Geometric isomers** (Sections 3.9 and 6.5): An old term for cis–trans isomers.

**Gibbs free-energy change,  $\Delta G$**  (Section 5.7): The total amount of free-energy change, both enthalpy and entropy, that occurs during a reaction. The Gibbs free-energy change for a reaction is given by the formula

$$\Delta G = \Delta H - T\Delta S$$

**Globular protein** (Section 27.13): Proteins that are coiled into compact, nearly spherical shapes. These proteins, which are generally water-soluble and mobile within the cell, are the structural class to which enzymes belong.

**Glycol** (Section 7.8): A 1,2-diol such as ethylene glycol,  $\text{HOCH}_2\text{CH}_2\text{OH}$ .

**Glycoside** (Section 24.8): A cyclic acetal formed by reaction of a sugar with another alcohol.

**Graft copolymer** (Section 31.8): A copolymer that consists of homopolymer chains grafted onto a different homopolymer backbone. Graft copolymers are prepared by X-ray irradiation of a homopolymer to generate radical sites along the chain, followed by addition of a second monomer.

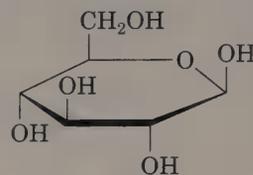
**Ground state** (Section 1.3): The most stable, lowest-energy electronic configuration of a molecule.

**Halohydrin** (Section 7.3): A 1,2-disubstituted haloalcohol such as is obtained on addition of HOBr to an alkene.

**Halonium ion** (Section 7.2): A species containing a positively charged, divalent halogen. Three-membered-ring bromonium ions are implicated as intermediates in the electrophilic addition of bromine to alkenes.

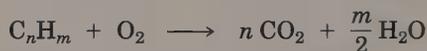
**Hammond postulate** (Section 6.12): A postulate stating that we can get a picture of what a given transition state looks like by looking at the structure of the nearest stable species. Exothermic reactions have transition states that resemble starting material, whereas endothermic reactions have transition states that resemble products.

**Haworth projection** (Section 24.5): A means of viewing stereochemistry in cyclic hemiacetal forms of sugars. Haworth projections are drawn so that the ring is flat and is viewed from an oblique angle with the hemiacetal oxygen at the upper right.



Haworth projection of glucose

**Heat of combustion** (Section 4.5): The amount of heat released when a compound is burned in a calorimeter according to the equation



**Heat of hydrogenation** (Section 6.7): The amount of heat released when a carbon-carbon double bond is hydrogenated. Comparison of heats of hydrogenation for different alkenes allows one to determine the stability of the different double bonds.

**Henderson-Hasselbalch equation** (Section 27.3): An equation for determining the extent of deprotonation of a weak acid at various pH's.

**Heterocycle** (Section 15.9 and Chapter 29): A cyclic molecule whose ring contains more than one kind of atom. For example, pyridine is a heterocycle that contains five carbon atoms and one nitrogen atom in its ring.

**Heterogenic bond formation** (Section 5.2): What occurs when one partner donates both electrons in forming a new bond. Polar reactions always involve heterogenic bond formation:

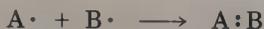


**Heterolytic bond breakage** (Section 5.2): The kind of bond breaking that occurs in polar reactions when one fragment leaves with both of the bonding electrons, as in the equation



**HOMO** (Sections 14.11 and 30.6): An acronym for highest occupied molecular orbital. The symmetries of the HOMO and LUMO are important in pericyclic reactions. (See LUMO.)

**Homogenic bond formation** (Section 5.2): What occurs in radical reactions when each partner donates one electron to the new bond:



**Homolytic bond breakage** (Section 5.2): The kind of bond breaking that occurs in radical reactions when each fragment leaves with one bonding electron according to the equation



**Homopolymer** (Section 31.8): A polymer made by chain-growth polymerization of a single monomer unit.

**Hückel's rule** (Section 15.7): A rule stating that monocyclic conjugated molecules having  $4n + 2$  pi electrons ( $n =$  an integer) show the unusual stability associated with aromaticity.

**Hybrid orbital** (Section 1.8): An orbital that is mathematically derived from a combination of ground-state ( $s$ ,  $p$ ,  $d$ ) atomic orbitals. Hybrid orbitals, such as the  $sp^3$ ,  $sp^2$ , and  $sp$  hybrids of carbon, are strongly directed and form stronger bonds than ground-state atomic orbitals do.

**Hydration** (Section 7.4): Addition of water to a molecule, such as occurs when alkenes are treated with strong sulfuric acid.

**Hydroboration** (Section 7.5): Addition of borane ( $BH_3$ ) or an alkylborane to an alkene. The resultant trialkylborane products are useful synthetic intermediates that can be oxidized to yield alcohols.

**Hydrogen bond** (Section 17.3): A weak (5 kcal/mol) attraction between a hydrogen atom bonded to an electronegative element and an electron lone pair on another atom. Hydrogen bonding plays an important role in determining the secondary structure of proteins and in stabilizing the DNA double helix.

**Hydrogenation** (Section 7.7): Addition of hydrogen to a double or triple bond to yield the saturated product.

**Hyperconjugation** (Section 6.7): A weak stabilizing interaction that results from overlap of a  $p$  orbital with a neighboring sigma bond. Hyperconjugation is important in stabilizing carbocations and in stabilizing substituted alkenes.

**Inductive effect** (Sections 2.4 and 16.5): The electron-attracting or electron-withdrawing effect that is transmitted through sigma bonds as the result of a nearby dipole. Electronegative elements have an electron-withdrawing inductive effect, whereas electropositive elements have an electron-donating inductive effect.

**Infrared spectroscopy** (Section 12.8): A kind of optical spectroscopy that uses infrared energy. IR spectroscopy is particularly useful in organic chemistry for determining the kinds of functional groups present in molecules.

**Initiator** (Section 5.4): A substance with an easily broken bond that is used to initiate radical chain reactions. For example, radical chlorination of alkanes is initiated when light energy breaks the weak chlorine–chlorine bond to form chlorine radicals.

**Intermediate** (Section 5.10): A species that is formed during the course of a multi-step reaction but is not the final product. Intermediates are more stable than transition states but may or may not be stable enough to isolate.

**Intramolecular, intermolecular** (Section 23.7): Reactions that occur within the same molecule are intramolecular, whereas reactions that occur between two molecules are intermolecular.

**Ion pair** (Section 11.8): A loose complex between two ions in solution. Ion pairs are implicated as intermediates in  $S_N1$  reactions in order to account for the partial retention of stereochemistry that is often observed.

**Ionic bond** (Section 1.5): A bond between two ions due to the electrical attraction of unlike charges. Ionic bonds are formed between strongly electronegative elements (such as the halogens) and strongly electropositive elements (such as the alkali metals).

**Ionization energy** (Section 1.5): The amount of energy required to remove an electron from an atom. Elements on the far right of the periodic table have high ionization energies, and elements on the far left of the periodic table have low ionization energies.

**Isoelectric point** (Section 27.3): The pH at which the number of positive charges and the number of negative charges on a protein or an amino acid are balanced.

**Isomers** (Section 3.2): Compounds that have the same molecular formula but have different structures.

**Isoprene rule** (Section 28.6): An observation to the effect that terpenoids appear to be made up of isoprene (2-methyl-1,3-butadiene) units connected in a head-to-tail fashion. Monoterpenes have two isoprene units, sesquiterpenes have three isoprene units, diterpenes have four isoprene units, and so on.

**Isotactic polymer** (Section 31.6): A chain-growth polymer in which all substituents on the polymer backbone have the same three-dimensional orientation.

**Kekulé structure** (Section 1.6): A representation of molecules in which a line between atoms is used to represent a bond. (See Line-bond structure.)

**Kinetic control** (Section 14.6): Reactions that follow the lowest activation-energy pathway are said to be kinetically controlled. The product formed in a kinetically controlled reaction is the one that is formed most rapidly, but is not necessarily the most stable. (See Thermodynamic control.)

**Kinetics** (Section 11.3): Referring to rates of reactions. Kinetics measurements can be extremely important in helping to determine reaction mechanisms.

**Leaving group** (Section 11.5): The group that is replaced in a substitution reaction. The best leaving groups in nucleophilic substitution reactions are those that form the most stable, least basic anions.

**Levorotatory** (Section 9.3): Used to describe an optically active substance that rotates the plane of polarization of plane-polarized light in a left-handed (counterclockwise) direction. (See Dextrorotatory.)

**Lewis acid** (Section 2.8): A substance having a vacant low-energy orbital that can accept an electron pair from a base. All electrophiles are Lewis acids, but transition metal salts such as  $AlCl_3$  and  $ZnCl_2$  are particularly good ones. (See Lewis base.)

**Lewis base** (Section 2.8): A substance that donates an electron lone pair to an acid. All nucleophiles are Lewis bases. (See Lewis acid.)

**Lewis structure** (Section 1.6): A representation of a molecule showing covalent bonds as a pair of electron dots between atoms.

**Line-bond structure** (Section 2.1): A representation of a molecule showing covalent bonds as lines between atoms. (See Kekulé structure.)

**Lipid** (Chapter 28): A naturally occurring substance isolated from cells and tissues by extraction with nonpolar solvents. Lipids belong to many different structural classes, including fats, terpenes, prostaglandins, and steroids.

**Lipophilic** (Section 28.2): Fat-loving. Long, nonpolar hydrocarbon chains tend to cluster together in polar solvents because of their lipophilic properties.

**Lone-pair electrons** (Section 1.12): Nonbonding electron pairs that occupy valence orbitals. It is the lone-pair electrons that are used by nucleophiles in their reactions with electrophiles.

**LUMO** (Sections 14.11 and 30.6): An acronym for lowest unoccupied molecular orbital. The symmetries of LUMO and HOMO are important in determining the stereochemistry of pericyclic reactions. (See HOMO.)

**Magnetic equivalence** (Section 13.9): Used to describe nuclei that have identical chemical and magnetic environments and that therefore absorb at the same place in the NMR spectrum. For example, the six hydrogens in benzene are magnetically equivalent, as are the six carbons.

**Markovnikov's rule** (Section 6.10): A guide for determining the regiochemistry (orientation) of electrophilic addition reactions. In the addition of HX to an alkene, the hydrogen atom becomes bonded to the alkene carbon that has fewer alkyl substituents. A modern statement of this same rule is that electrophilic addition reactions proceed via the most stable carbocation intermediate.

**Mechanism** (Section 5.2): A complete description of how a reaction occurs. A mechanism must account for all starting materials and all products, and must describe the details of each individual step in the overall reaction process.

**Meso compound** (Section 9.8): A meso compound is one that contains stereogenic centers but is nevertheless achiral by virtue of a symmetry plane. For example, (2*R*,3*S*)-butanediol has two chiral carbon atoms but is achiral because of a symmetry plane between carbons 2 and 3.

**Micelle** (Section 28.2): A spherical cluster of soap-like molecules that aggregate in aqueous solution. The ionic heads of the molecules lie on the outside where they are solvated by water, and the organic tails bunch together on the inside of the micelle.

**Mobile phase** (Section 12.1): The solvent (either gas or liquid) used in chromatography to move material along the solid adsorbent phase. (See Chromatography.)

**Molecular formula** (Section 2.9): An expression of the total numbers of each kind of atom present in a molecule. The molecular formula must be a whole-number multiple of the empirical formula.

**Molecular ion** (Section 12.5): The cation produced in the mass spectrometer by loss of an electron from the parent molecule. The mass of the molecular ion corresponds to the molecular weight of the sample.

**Molecular orbital** (Section 1.7): An orbital that is the property of the entire molecule rather than of an individual atom. Molecular orbitals result from overlap of two or more atomic orbitals when bonds are formed and may be either bonding, nonbonding, or antibonding. Bonding molecular orbitals are lower in energy than the starting atomic orbitals, nonbonding MO's are equal in energy to the starting orbitals, and antibonding orbitals are higher in energy.

**Monomer** (Section 31.2): The simple starting units from which polymers are made.

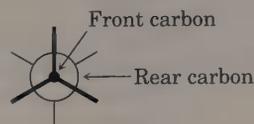
**Multiplet** (Section 13.6): A symmetrical pattern of peaks in an NMR spectrum that arises by spin-spin splitting of a single absorption because of coupling between neighboring magnetic nuclei.

**Mutarotation** (Section 24.6): The spontaneous change in optical rotation observed when a pure anomer of a sugar is dissolved in water. Mutarotation is caused by the reversible opening and closing of the acetal linkage, which yields an equilibrium mixture of anomers.

**$n + 1$  rule** (Sections 13.6 and 13.12): A carbon bonded to  $n$  hydrogens shows  $n + 1$  peaks in its spin-coupled  $^{13}\text{C}$  NMR spectrum, and a hydrogen with  $n$  other hydrogens on neighboring carbons shows  $n + 1$  peaks in its  $^1\text{H}$  NMR spectrum.

**Neighboring-group effect** (Section 24.8): The effect on a reaction of a nearby functional group.

**Newman projection** (Section 4.1): A means of indicating stereochemical relationships between substituent groups on neighboring carbons. The carbon-carbon bond is viewed end-on, and the carbons are indicated by a circle. Bonds radiating from the center of the circle are attached to the front carbon, and bonds radiating from the edge of the circle are attached to the rear carbon.



**Nitrogen rule** (Sections 12.6 and 25.11): A rule stating that compounds having an odd number of nitrogens give rise to an odd-numbered molecular ion in the mass spectrum. Conversely, compounds with an even number of nitrogens give rise to even-numbered molecular ions.

**Node** (Sections 1.2 and 14.3): The surface of zero electron density between lobes of orbitals. For example, a  $p$  orbital has a nodal plane passing through the center of the nucleus, perpendicular to the line of the orbital.

**Normal alkane** (Section 3.2): A straight-chain alkane, as opposed to a branched alkane. Normal alkanes are denoted by the suffix  $n$ , as in  $n\text{-C}_4\text{H}_{10}$  ( $n$ -butane).

**Nuclear magnetic resonance, NMR** (Chapter 13): A spectroscopic technique that provides information about the carbon-hydrogen framework of a molecule. NMR works by detecting the energy absorption accompanying the transition between nuclear spin states that occurs when a molecule is placed in a strong magnetic field and irradiated with radio-frequency waves. Different nuclei within a molecule are in slightly different magnetic environments and therefore show absorptions at slightly different frequencies.

**Nucleophile** (Section 5.5): A “nucleus-lover,” or species that donates an electron pair to an electrophile in a polar bond-forming reaction. Nucleophiles are also Lewis bases. (See Electrophile.)

**Nucleoside** (Section 29.9): A nucleic acid constituent, consisting of a sugar residue bonded to a heterocyclic purine or pyrimidine base.

**Nucleotide** (Section 29.9): A nucleic acid constituent, consisting of a sugar residue bonded both to a heterocyclic purine or pyrimidine base and to a phosphoric acid. Nucleotides are the monomer units from which DNA and RNA are constructed.

**Off-resonance mode** (Section 13.6): A mode of  $^{13}\text{C}$  NMR spectrometer operation that allows for the observation of spin-spin splitting between carbons and their attached hydrogens. Carbons bonded to one hydrogen show a doublet; carbons attached to two hydrogens show a triplet; and carbons attached to three hydrogens show a quartet in the off-resonance NMR.

**Olefin**: An alternative name for an alkene.

**Optical isomers** (Section 9.5): Enantiomers. Optical isomers are isomers that have a mirror-image relationship.

**Optically active** (Section 9.3): A substance that rotates the plane of polarization of plane-polarized light. Note that an optically active sample must contain chiral molecules, but that all samples with chiral molecules are not optically active. Thus, a racemic sample is optically inactive even though the individual molecules are chiral. (See Chiral.)

**Orbital** (Section 1.2): The volume of space in which an electron is most likely to be found. Orbitals are described mathematically by wave functions, which define the behavior of electrons around nuclei.

**Ozonide** (Section 7.9): The product formed by addition of ozone to a carbon-carbon double bond. Ozonides are usually treated with a reducing agent, such as zinc in acetic acid, to produce carbonyl compounds.

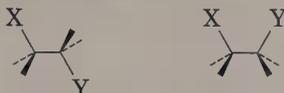
**Paraffins** (Section 3.6): A trivial name for alkanes.

**Pauli exclusion principle** (Section 1.3): A statement of the fact that no more than two electrons can occupy the same orbital and that those two must have spins of opposite sign.

**Peptides** (Section 27.6): Amino-acid polymers in which the individual amino acid residues are linked by amide bonds. (See Proteins.)

**Pericyclic reaction** (Chapter 30): A reaction that occurs by a concerted reorganization of bonding electrons in a cyclic transition state.

**Periplanar** (Section 11.11): A conformation in which bonds to neighboring atoms have a parallel arrangement. In an eclipsed conformation, the neighboring bonds are syn periplanar; in a staggered conformation, the bonds are anti periplanar.



Anti periplanar

Syn periplanar

**Peroxide** (Section 18.2): A molecule containing an oxygen-oxygen bond functional group, R-O-O-R' or R-O-O-H. The "peroxides" present as explosive impurities in ether solvents are usually of the latter type. Since the oxygen-oxygen bond is weak and easily broken, peroxides are often used to initiate radical chain reactions.

**Phase-transfer catalysts** (Section 25.9): Agents that cause the transfer of ionic reagents between phases, thus catalyzing reactions. Tetraalkylammonium salts,  $R_4N^+ X^-$ , are often used to transport inorganic anions from the aqueous phase to the organic phase where the desired reaction then occurs. For example, permanganate ion is solubilized in benzene in the presence of tetraalkylammonium ions.

**Phospholipid** (Section 28.3): Lipids that contain a phosphate residue. For example, phosphoglycerides contain a glycerol backbone linked to two fatty acids and a phosphoric acid.

**Pi bond** (Section 1.7): The covalent bond formed by sideways overlap of atomic orbitals. For example, carbon-carbon double bonds contain a pi bond formed by sideways overlap of two *p* orbitals.

**Plane of symmetry** (Section 9.2): An imaginary plane that bisects a molecule such that one half of the molecule is the mirror image of the other half. Molecules containing a plane of symmetry are achiral.

**Plane-polarized light** (Section 9.3): Ordinary light that has its electric vectors in a single plane rather than in random planes. The plane of polarization is rotated when the light is passed through a solution of a chiral substance.

**Polar aprotic solvent** (Section 11.5): Polar solvent that can't function as a hydrogen ion donor. Polar aprotic solvents such as dimethyl sulfoxide (DMSO), hexamethylphosphoramide (HMPA), and dimethylformamide (DMF) are particularly useful in  $S_N2$  reactions because of their ability to solvate cations.

**Polar reaction** (Section 5.2): A reaction in which bonds are made when a nucleophile donates two electrons to an electrophile and in which bonds are broken when one fragment leaves with both electrons from the bond. Polar reactions are the most common class of reactions. (See Heterogenic bond formation; Heterolytic bond breakage.)

**Polarity** (Sections 2.4 and 5.5): The unsymmetrical distribution of electrons in molecules that results when one atom attracts electrons more strongly than another.

**Polarizability** (Section 5.5): The measure of the change in a molecule's electron distribution in response to changing electric interactions with solvents or ionic reagents.

**Polymer** (Chapter 31): A large molecule made up of repeating smaller units. For example, polyethylene is a synthetic polymer made from repeating ethylene units, and DNA is a biopolymer made of repeating deoxyribonucleotide units.

**Primary, secondary, tertiary, quaternary** (Section 3.4): Terms used to describe the substitution pattern at a specific site. A primary site has one organic substituent attached to it, a secondary site has two organic substituents, a tertiary site has three, and a quaternary site has four.

	<i>Carbon</i>	<i>Hydrogen</i>	<i>Alcohol</i>	<i>Amine</i>
Primary	$RCH_3$	$RCH_3$	$RCH_2OH$	$RNH_2$
Secondary	$R_2CH_2$	$R_2CH_2$	$R_2CHOH$	$R_2NH$
Tertiary	$R_3CH$	$R_3CH$	$R_3COH$	$R_3N$
Quaternary	$R_4C$			

R = any organic substituent

**Primary structure** (Section 27.14): The amino acid sequence in a protein. (See Secondary structure; Tertiary structure.)

**Principle of maximum overlap** (Section 1.7): The strongest bonds are formed when overlap between orbitals is greatest.

**Propagation step** (Section 5.4): The step or series of steps in a radical chain reaction that carry on the chain. The propagation steps must yield both product and a reactive intermediate to carry on the chain.

**Prostaglandin** (Section 28.5): A member of the class of lipids with the general carbon skeleton.



Prostaglandins are present in nearly all body tissues and fluids, where they serve a large number of important hormonal functions.

**Protecting group** (Section 17.10): A group that is introduced to protect a sensitive functional group toward reaction elsewhere in the molecule. After serving its protective function, the group is then removed. For example, ketones and aldehydes are often protected as acetals by reaction with ethylene glycol, and alcohols are often protected as trimethylsilyl ethers.

**Protein** (Section 27.13): A large peptide containing fifty or more amino acid residues. Proteins serve both as structural materials (hair, horns, fingernails) and as enzymes that control an organism's chemistry. (*See* Peptides.)

**Protic solvent** (Section 11.9): A solvent such as water or alcohol that can serve as a proton donor. Protic solvents are particularly good at stabilizing anions by hydrogen bonding, thereby lowering their reactivity. (*See* Polar aprotic solvent.)

**Proton noise-decoupled mode** (Section 13.4): The most common manner of  $^{13}\text{C}$  NMR spectrometer operation, in which all nonequivalent carbon atoms in the sample show a single, unsplit resonance. Operating the spectrometer in this mode allows one to count the number of chemically different carbon atoms present in the sample molecule.

**Quartet** (Section 13.6): A set of four peaks in the NMR, caused by spin-spin splitting of a signal by three adjacent nuclear spins.

**Quaternary** (Section 3.4): *See* Primary.

**Quaternary structure** (Section 27.14): The highest level of protein structure, involving a specific aggregation of individual proteins into a larger cluster.

***R,S* convention** (Section 9.6): A method for defining the absolute configuration around stereogenic centers. The Cahn-Ingold-Prelog sequence rules are used to assign relative priorities to the four substituents on the stereogenic center, and the center is oriented such that the group of lowest (fourth) priority faces directly away from the viewer. If the three remaining substituents have a right-handed or clockwise relationship in going from first to second to third priority, then the stereogenic center is denoted *R* (rectus, right). If the three remaining substituents have a left-handed or counterclockwise relationship, the stereogenic center is denoted *S* (sinister, left). (*See* Sequence rules.)

**Racemic mixture** (Section 9.10): A mixture consisting of equal parts (+) and (−) enantiomers of a chiral substance. Even though the individual molecules are chiral, racemic mixtures are optically inactive.

**Racemization** (Section 9.11): The process whereby one enantiomer of a chiral molecule becomes converted into a 50:50 mixture of enantiomers, thus losing its optical activity. For example, this might happen during an  $\text{S}_{\text{N}}1$  reaction of a chiral alkyl halide.

**Radical** (Section 5.2): When used in organic nomenclature, the word *radical* refers to a part of a molecule that appears in its name—for example, the “phenyl” in phenyl acetate. Chemically, however, a radical is a species that has an odd number of electrons, such as the chlorine radical,  $\text{Cl}\cdot$ .

**Radical reaction** (Section 5.2): A reaction in which bonds are made by donation of one electron from each of two reagents and in which bonds are broken when each fragment leaves with one electron. (*See* Homogenic bond formation; Homolytic bond breakage.)

**Rate-limiting step** (Section 11.7): The slowest step in a multistep reaction sequence. The rate-limiting step acts as a kind of bottleneck in multistep reactions and is observed by kinetics measurements.

**Reaction energy diagram** (Section 5.9): A pictorial representation of the course of a reaction, in which potential energy is plotted as a function of reaction progress. Starting materials, transition states, intermediates, and final products are all represented, and their appropriate energy levels are indicated.

**Reducing sugar** (Section 24.8): Any sugar that reduces silver ion in the Tollens test or cupric ion in the Fehling or Benedict tests. All sugars that are aldehydes or

can be readily converted into aldehydes are reducing. Glycosides, however, are not reducing sugars.

**Refining** (Section 3.5): The process by which petroleum is converted into gasoline and other useful products.

**Regiochemistry** (Section 6.10): A term describing the orientation of a reaction that occurs on an unsymmetrical substrate. Markovnikov's rule, for example, predicts the regiochemistry of electrophilic addition reactions.

**Regiospecific** (Section 6.10): A term describing a reaction that occurs with a specific regiochemistry to give a single product rather than a mixture of products.

**Replication** (Section 29.13): The process by which double-stranded DNA uncoils and is replicated to produce two new copies.

**Resolution** (Section 25.5): The process by which a racemic mixture is separated into its two pure enantiomers. For example, a racemic carboxylic acid might be converted by reaction with a chiral amine base into a diastereomeric mixture of salts, which could be separated by fractional crystallization. Regeneration of the free acids would then yield the two pure enantiomeric acids.

**Resonance effect** (Section 16.5): The effect by which substituents donate or withdraw electron density through orbital overlap with neighboring pi bonds. For example, an oxygen or nitrogen substituent donates electron density to an aromatic ring by overlap of the O or N orbital with the aromatic ring *p* orbitals. A carbonyl substituent, however, withdraws electron density from an aromatic ring by *p* orbital overlap. These effects are particularly important in determining whether a given group is meta-directing or ortho- and para-directing in electrophilic aromatic substitution reactions.

**Resonance hybrid** (Section 10.6): A molecule, such as benzene, that can't be represented adequately by a single Kekulé structure but must instead be considered as an average of two or more resonance structures. The resonance structures themselves differ only in the positions of their electrons, not their nuclei.

**Ring current** (Section 15.12): The circulation of pi electrons induced in aromatic rings by an external magnetic field. This effect accounts for the pronounced downfield shift of aromatic-ring protons in the  $^1\text{H}$  NMR.

**Ring-flip** (Section 4.11): The molecular motion that converts one chair conformation of cyclohexane into another chair conformation. The effect of a ring-flip is to convert an axial substituent into an equatorial substituent.

**RNA** (Section 29.9): Ribonucleic acid, the biopolymer found in cells that serves to transcribe the genetic information found in DNA and uses that information to direct the synthesis of proteins.

**Saccharide** (Section 24.1): A sugar.

**Saponification** (Section 21.7): An old term for the base-induced hydrolysis of an ester to yield a carboxylic acid salt.

**Saturated** (Section 3.2): A saturated molecule is one that has only single bonds and thus can't undergo addition reactions. Alkanes, for example, are saturated, but alkenes are unsaturated.

**Sawhorse structure** (Section 4.1): A stereochemical manner of representation that portrays a molecule using a stick drawing and gives a perspective view of the conformation around single bonds.



Sawhorse structure

**Second-order reaction** (Section 11.3): A reaction whose rate-limiting step is bimolecular and whose kinetics are therefore dependent on the concentration of two reagents.

**Secondary** (Section 3.4): *See* Primary.

**Secondary structure** (Section 27.14): The level of protein substructure that involves organization of chain sections into ordered arrangements such as  $\beta$ -pleated sheets or  $\alpha$ -helices.

**Sequence rules** (Sections 6.6 and 9.6): A series of rules devised by Cahn, Ingold, and Prelog for assigning relative priorities to substituent groups on a double-bond carbon atom or on a stereogenic center. Once priorities have been established, *E,Z* double-bond geometry and *R,S* configurational assignments can be made. (*See Entgegen*; *R,S* convention; *Zusammen*.)

**Shielding** (Section 13.2): An effect observed in NMR that causes a nucleus to absorb toward the right (upfield) side of the chart. Shielding is caused by donation of electron density to the nucleus. (*See* Deshielding.)

**Sigma bond** (Section 1.7): A covalent bond formed by head-on overlap of atomic orbitals.

**Sigmatropic reaction** (Sections 30.1 and 30.10): A pericyclic reaction that involves the migration of a group from one end of a pi electron system to the other. For example, the [1.5] sigmatropic rearrangement of a hydrogen atom in cyclopentadiene is such a reaction.

**Soap** (Section 28.2): The mixture of long-chain fatty acid salts obtained on base hydrolysis of animal fat.

**Solid-phase synthesis** (Section 27.12): A technique of synthesis whereby the starting material is covalently bound to a solid polymer bead and reactions are carried out on the bound substrate. After the desired transformations have been effected, the product is cleaved from the polymer and is isolated. This technique is particularly useful in peptide synthesis (Merrifield method).

***sp* orbital** (Section 1.11): A hybrid orbital mathematically derived from the combination of an *s* and a *p* atomic orbital. The two *sp* orbitals that result from hybridization are oriented at an angle of 180° to each other.

***sp*<sup>2</sup> orbital** (Section 1.10): A hybrid orbital mathematically derived by combination of an *s* atomic orbital with two *p* atomic orbitals. The three *sp*<sup>2</sup> hybrid orbitals that result lie in a plane at angles of 120° to each other.

***sp*<sup>3</sup> orbital** (Section 1.8): A hybrid orbital mathematically derived by combination of an *s* atomic orbital with three *p* atomic orbitals. The four *sp*<sup>3</sup> hybrid orbitals that result are directed toward the corners of a tetrahedron at angles of 109° to each other.

**Specific rotation,  $[\alpha]_D$**  (Section 9.4): The specific rotation of a chiral compound is a physical constant that is defined by the equation

$$[\alpha]_D = \frac{\text{Observed rotation}}{\text{Path length} \times \text{Concentration}} = \frac{\alpha}{l \times C}$$

where the path length of the sample solution is expressed in decimeters and the concentration of the sample solution is expressed in g/mL.

**Spin-spin splitting** (Section 13.6): The splitting of an NMR signal into a multiplet caused by an interaction between nearby magnetic nuclei whose spins are coupled. The magnitude of spin-spin splitting is given by the coupling constant, *J*.

**Staggered conformation** (Section 4.1): The three-dimensional arrangement of atoms around a carbon-carbon single bond in which the bonds on one carbon exactly bisect the bond angles on the second carbon as viewed end-on. (See Eclipsed conformation.)



Staggered conformation

**Stationary phase** (Section 12.1): The solid support used in chromatography. The molecules to be chromatographically separated adsorb to the stationary phase and are moved along by the mobile phase. Silica gel (hydrated  $\text{SiO}_2$ ) and alumina ( $\text{Al}_2\text{O}_3$ ) are often used as stationary phases in column chromatography of organic mixtures. (See Mobile phase.)

**Step-growth polymer** (Section 31.1): A polymer produced by a series of polar reactions between two difunctional monomers. The polymer normally has the two monomer units in alternating order and usually has other atoms in addition to carbon in the polymer backbone. Nylon, a polyamide produced by reaction between a diacid and a diamine, is an example of such a polymer.

**Stereochemistry** (Chapters 4 and 9): The branch of chemistry concerned with the three-dimensional arrangement of atoms in molecules.

**Stereogenic center** (Section 9.2): An atom (usually carbon) that is bonded to four different groups and is therefore chiral. (See Chiral.)

**Stereoisomers** (Section 3.9): Isomers that have their atoms connected in the same order but have different three-dimensional arrangements. The term *stereoisomer* includes both enantiomers and diastereomers but does not include constitutional isomers.

**Stereospecific** (Section 7.11): A term indicating that only a single stereoisomer is produced in a given reaction, rather than a mixture.

**Steric strain** (Sections 4.3 and 4.12): The strain imposed on a molecule when two groups are too close together and try to occupy the same space. Steric strain is responsible both for the greater stability of trans versus cis alkenes and for the greater stability of equatorially substituted versus axially substituted cyclohexanes.

**Steroid** (Section 28.8): A lipid whose structure is based on the tetracyclic carbon skeleton:



Steroids occur in both plants and animals and have a variety of important hormonal functions.

**Suprafacial** (Section 30.8): A word used to describe the geometry of pericyclic reactions. Suprafacial reactions take place on the same side of the two ends of a pi electron system. (See Antarafacial.)

**Symmetry-allowed, symmetry-disallowed** (Section 30.4): A symmetry-allowed reaction is a pericyclic process that has a favorable orbital symmetry for reaction through a concerted pathway. A symmetry-disallowed reaction is one that does not have favorable orbital symmetry for reaction through a concerted pathway.

**Syn stereochemistry** (Section 7.2): A syn addition reaction is one in which the two ends of the double bond are attacked from the same side. For example,  $\text{OsO}_4$  induced hydroxylation of cyclohexene yields *cis*-1,2-cyclohexanediol, the product of syn addition. A syn elimination is one in which the two groups leave from the same side of the molecule.

**Syndiotactic polymer** (Section 31.6): A chain-growth polymer in which the substituents on the polymer backbone have a regular alternating stereochemistry.

**Tautomers** (Sections 8.5 and 22.1): Isomers that are rapidly interconverted. For example, enols and ketones are tautomers since they are rapidly interconverted on treatment with either acid or base catalysts.

**Terpenes** (Section 28.6): Lipids that are formally derived by head-to-tail polymerization of isoprene units. (See Isoprene rule.)

**Tertiary** (Section 3.4): See Primary.

**Tertiary structure** (Section 27.14): The level of protein structure that involves the manner in which the entire protein chain is folded into a specific three-dimensional arrangement.

**Thermodynamic control** (Section 14.6): Equilibrium reactions that yield the lowest-energy, most stable product are said to be thermodynamically controlled. Although most stable, the product of a thermodynamically controlled reaction is not necessarily formed fastest. (See Kinetic control.)

**Thermoplastic** (Section 31.13): A polymer that is hard at room temperature but that becomes soft and pliable when heated. Thermoplastics are used for the manufacture of a variety of molded objects.

**Thermosetting resin** (Section 31.13): A polymer that is highly cross-linked and sets into a hard, insoluble mass when heated. Bakelite is the best-known example of such a polymer.

**Torsional strain** (Section 4.1): The strain in a molecule caused by electron repulsion between eclipsed bonds. Torsional strain plays a major role in destabilizing boat cyclohexane relative to chair cyclohexane. (See Eclipsing strain.)

**Transcription** (Section 29.14): The process by which the genetic information encoded in DNA is read and used to synthesize RNA in the nucleus of the cell. A small portion of double-stranded DNA uncoils, and complementary ribonucleotides line up in the correct sequence for RNA synthesis.

**Transition state** (Section 5.9): An imaginary activated complex between reagents, representing the highest energy point on a reaction curve. Transition states are unstable complexes that can't be isolated.

**Translation** (Section 29.15): The process by which the genetic information transcribed from DNA onto mRNA is read by tRNA and used to direct protein synthesis.

**Tree diagram** (Section 13.13): A diagram used in NMR to help sort out the complicated splitting patterns that can arise from multiple couplings.

**Triacylglycerol** (Section 28.1): Lipids such as animal fat and vegetable oil consisting chemically of triesters of glycerol with long-chain fatty acids.

**Triplet** (Section 13.6): A symmetrical three-line splitting pattern observed in the  $^1\text{H}$  NMR when a proton has two equivalent neighbor protons or in the  $^{13}\text{C}$  NMR when a carbon is bonded to two hydrogens.

**Ultraviolet (UV) spectroscopy** (Section 14.10): An optical spectroscopy employing ultraviolet irradiation. UV spectroscopy provides structural information about the extent of pi electron conjugation in organic molecules.

**Unimolecular reaction** (Section 11.7): A reaction that occurs by spontaneous transformation of the starting material without the intervention of other reagents. For example, the dissociation of a tertiary alkyl halide in the  $\text{S}_{\text{N}}1$  reaction is a unimolecular process.

**Unsaturated** (Section 6.2): An unsaturated molecule is one that has multiple bonds and can undergo addition reactions. Alkenes and alkynes, for example, are unsaturated. (See Saturated.)

**Upfield** (Section 13.3): Used to refer to the right-hand portion of the NMR chart. (See Shielding.)

**Van der Waals forces** (Section 3.6): The attractive forces between molecules that are caused by dipole–dipole interactions. Van der Waals forces are one of the primary forces responsible for holding molecules together in the liquid and solid states.

**Vicinal** (Section 8.3): A term used to refer to a 1,2-disubstitution pattern. For example, 1,2-dibromoethane is a vicinal dibromide.

**Vinylic** (Section 8.4): A term that refers to a substituent at a double–bond carbon atom. For example, chloroethylene is a vinylic chloride, and enols are vinylic alcohols.

**Vulcanization** (Section 31.7): A process for hardening rubber by heating in the presence of elemental sulfur. The sulfur functions by forming cross-links between polymer chains.

**Wave function** (Section 1.1): The mathematical expression that defines the behavior of an electron. The square of the wave function is the probability function that defines the shapes of orbitals.

**Wave number** (Section 12.9): The wave number is the reciprocal of the wavelength in centimeters. Thus, wave numbers are expressed in  $\text{cm}^{-1}$ .

**Wavelength** (Section 12.8): The length of a wave from peak to peak. The wavelength of electromagnetic radiation is inversely proportional to frequency and inversely proportional to energy. (See Frequency.)

**Ylide** (Section 19.16): A neutral dipolar molecule in which the positive and negative charges are adjacent. The phosphoranes used in Wittig reactions are ylides.

**Zaitsev's rule** (Section 11.10): A rule stating that E2 elimination reactions normally yield the more highly substituted alkene as major product.

**Zusammen (Z)** (Section 6.6): A term used to describe the stereochemistry of a carbon–carbon double bond. The two groups on each carbon are assigned priorities according to the Cahn-Ingold-Prelog sequence rules, and the two carbons are compared. If the high-priority groups on each carbon are on the same side of the double bond, the bond has Z geometry. (See *Entgegen*; Sequence rules.)

**Zwitterion** (Section 27.2): A neutral dipolar molecule in which the positive and negative charges are not adjacent. For example, amino acids exist as zwitterions,  $\text{H}_3\text{N}^+ - \text{CHR} - \text{COO}^-$ . Zwitterions are also called betaines.

# Appendix D

## Answers to Selected In-Text Problems

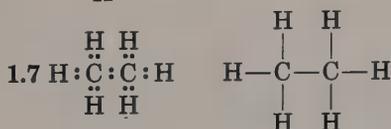
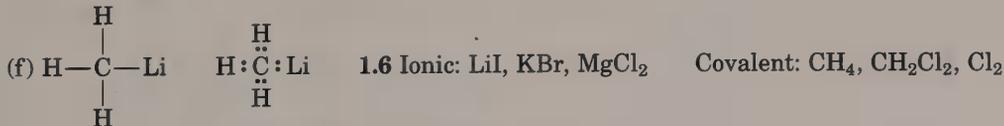
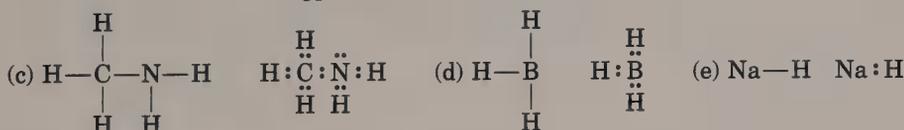
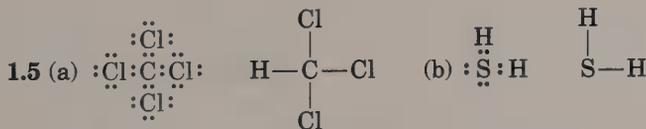
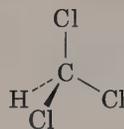
The following answers are meant only as a quick check. Full answers and explanations for all problems are provided in the accompanying *Study Guide and Solutions Manual*.

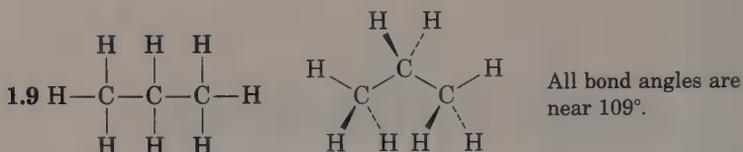
### CHAPTER 1

1.1 (a) 1 (b) 3 (c) 8 1.2 (a) Boron:  $1s^2 2s^2 2p$  (b) Phosphorus:  $1s^2 2s^2 2p^6 3s^2 3p$

(c) Oxygen:  $1s^2 2s^2 2p^4$  (d) Chlorine:  $1s^2 2s^2 2p^6 3s^2 3p^5$  1.3

1.4 (a)  $\text{CCl}_4$  (b)  $\text{AlH}_3$  (c)  $\text{CH}_2\text{Cl}_2$  (d)  $\text{SiF}_4$  (e)  $\text{CH}_3\text{NH}_2$

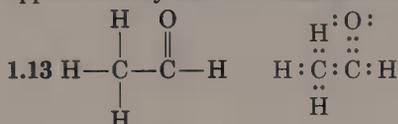




1.10 C<sub>2</sub>H<sub>7</sub> has too many hydrogens.

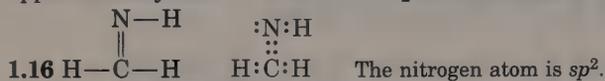
1.11 The CH<sub>3</sub> carbon is sp<sup>3</sup>; double-bond carbons are sp<sup>2</sup>. The C-C-C bond angle is approximately 120°.

1.12 All carbons are sp<sup>2</sup>; all bond angles are near 120°.



1.14 The CH<sub>3</sub> carbon is sp<sup>3</sup>; triple-bond carbons are sp. The C-C-C bond angle is approximately 180°.

1.15 The N<sub>2</sub> molecule has a triple bond, N≡N.



1.17 All are sp<sup>3</sup> hybridized and have tetrahedral geometry.

## CHAPTER 2

2.1 (a) C<sub>5</sub>H<sub>5</sub>N (b) C<sub>6</sub>H<sub>10</sub>O (c) C<sub>8</sub>H<sub>7</sub>N

2.2 (a) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>3</sub> (b) CH<sub>3</sub>CH<sub>2</sub>NH<sub>2</sub> (c) CH<sub>3</sub>CH<sub>2</sub>CHO (d) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Cl

2.4 For sulfur:

$$\text{F.C.} = 6 \text{ valence electrons} - \frac{6 \text{ bonding electrons}}{2} - 2 \text{ nonbonding electrons} = +1$$

For oxygen:

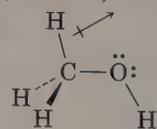
$$\text{F.C.} = 6 \text{ valence electrons} - \frac{2 \text{ bonding electrons}}{2} - 6 \text{ nonbonding electrons} = -1$$

2.5 (a) C = 0, central N = +1, terminal N = -1 (b) C = 0, N = +1, O = -1

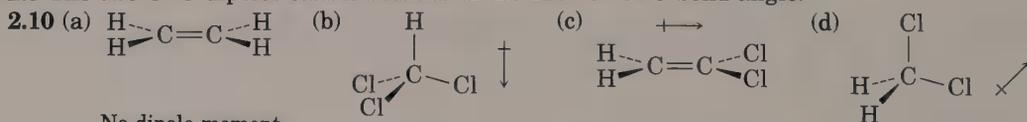
(c) N = -1, C = +1

2.6 (a) H (b) Br (c) Cl (d) C

2.7 (a) C is δ<sup>+</sup>, Br is δ<sup>-</sup> (b) C is δ<sup>+</sup>, N is δ<sup>-</sup> (c) Li is δ<sup>+</sup>, C is δ<sup>-</sup> (d) H is δ<sup>+</sup>, N is δ<sup>-</sup>

(e) C is δ<sup>+</sup>, O is δ<sup>-</sup> (f) Mg is δ<sup>+</sup>, C is δ<sup>-</sup> (g) C is δ<sup>+</sup>, F is δ<sup>-</sup> 2.8 

2.9 The two C-O dipoles cancel because of the 180° O=C=O bond angle.



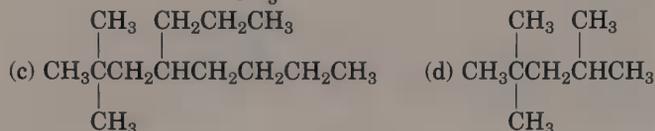
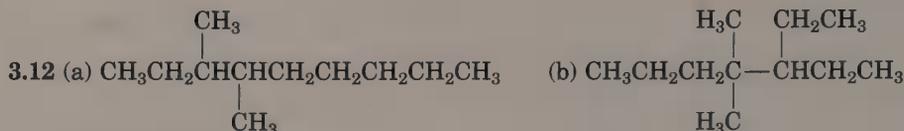
No dipole moment

2.11 Picric acid is stronger. 2.12 Water is a stronger acid.

2.13 Neither reaction will take place. 2.14 Reaction will take place. 2.15 6.3 × 10<sup>10</sup>

2.16 pH = 2.25 2.17 Lewis acids: (c), (d), (e) Lewis bases: (a), (b), (f)





3.13 (a) 2-Methylhexane (b) 4,5-Dimethyloctane (c) 3-Ethyl-2,2-dimethylpentane

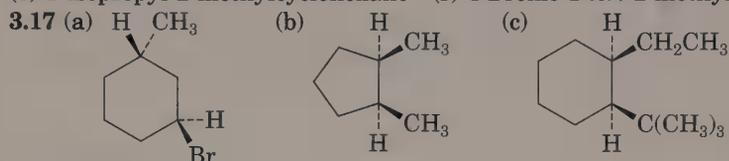
(d) 3,4-Dimethylheptane (e) 2,3-Dimethylhexane (f) 3,3-Dimethylpentane

3.14 Pentyl, 1-methylbutyl, 1-ethylpropyl, 2-methylbutyl, 3-methylbutyl, 1,1-dimethylpropyl, 1,2-dimethylpropyl, 2,2-dimethylpropyl

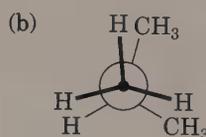
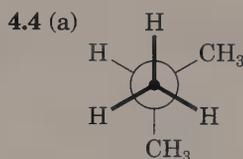
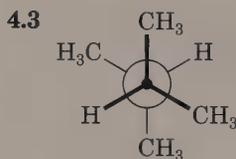
3.15 (a) 1,4-Dimethylcyclohexane (b) 1-Methyl-3-propylcyclopentane

(c) 3-Cyclobutylpentane (d) 1-Bromo-4-ethylcyclohexane

(e) 1-Isopropyl-2-methylcyclohexane (f) 4-Bromo-1-*tert*-2-methylcycloheptane



## CHAPTER 4



4.5 Six H-H interactions; 22% of strain.

4.6 The *trans* isomer is more stable and has lower heat of combustion.

4.7 Cyclopropane

4.8 Methyl groups in *cis*-1,2-dimethylcyclobutane and in *trans*-1,3-dimethylcyclobutane have steric interactions.

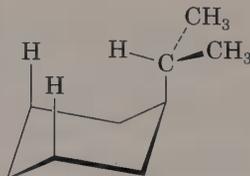
4.9



4.10 Ten H-H eclipsing interactions; 10 kcal/mol.

4.12–4.13 Axial and equatorial positions alternate around the ring on each side.

4.15 Hydrogen of isopropyl group points in toward ring:



4.16 Cyano group points straight up. 4.17 Equatorial > 99.9%

4.18 (a) 0.5 kcal/mol (b) 1.8 kcal/mol (c) 0.5 kcal/mol (d) 1.9 kcal/mol

4.20 *trans*-Decalin is more stable by 2.7 kcal/mol.

## CHAPTER 5

5.1 (a) Substitution (b) Elimination (c) Addition

5.2 Reaction proceeds until stopped by random termination steps.

5.3 1-Chloro-2-methylpentane, 2-chloro-2-methylpentane, 3-chloro-2-methylpentane, 2-chloro-4-methylpentane, 1-chloro-4-methylpentane

5.4 Pentane has three types of hydrogens; neopentane has only one.

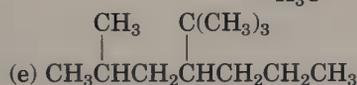
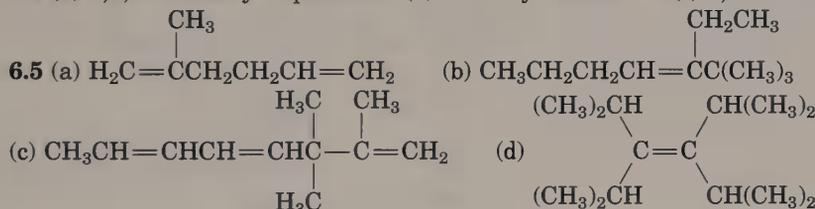
5.6 Electrophiles:  $H^+$ ,  $Br^+$ ,  $Mg^{2+}$ ,  $CO_2$ ; nucleophiles:  $HO^-$ ,  $NH_3$  5.7 Bromocyclohexane5.8 Negative  $\Delta G^\circ$  is more favored. 5.9 Larger  $K_{eq}$  is more exothermic.5.10  $\Delta G^\circ = -4.09$  kcal/mol, 0 kcal/mol,  $+4.09$  kcal/mol;  $K_{eq} = 2.2 \times 10^7$ , 0,  $4.6 \times 10^{-8}$ 5.11 (a)  $\Delta H^\circ = +58$  kcal/mol (b)  $\Delta H^\circ = +1$  kcal/mol (c)  $\Delta H^\circ = -26$  kcal/mol(d)  $\Delta H^\circ = -25$  kcal/mol 5.12 (a)  $\Delta H^\circ = -7$  kcal/mol (b)  $\Delta H^\circ = +4$  kcal/mol5.13 Lower  $\Delta G^\ddagger$  is faster. Can't predict  $K_{eq}$ .

## CHAPTER 6

6.1 (a) 2 (b) 3 (c) 3 (d) 5 (e) 13 6.2 (a) 1 (b) 2 (c) 2

6.3 (a) 5 (b) 5 (c) 3 (d) 1 (e) 6 (f) 5

6.4 (a) 3,4,4-Trimethyl-1-pentene (b) 3-Methyl-3-hexene (c) 4,7-Dimethyl-2,5-octadiene



6.6 (a) 1,2-Dimethylcyclohexene (b) 4,4-Dimethylcycloheptene (c) 3-Isopropylcyclopentane

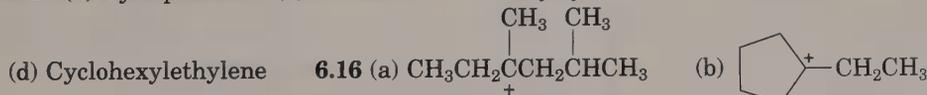
6.7 Compounds (c), (e), and (f) have cis-trans isomers.

6.8 *trans*-Cyclohexene is too strained.6.9 (a)  $-Br$  (b)  $-Br$  (c)  $-CH_2CH_3$  (d)  $-OH$  (e)  $-CH_2OH$  (f)  $-CH=O$ 6.10 (a)  $-Cl$ ,  $-OH$ ,  $-CH_3$ ,  $-H$  (b)  $-CH_2OH$ ,  $-CH=CH_2$ ,  $-CH_2CH_3$ ,  $-CH_3$ (c)  $-COOH$ ,  $-CH_2OH$ ,  $-CN$ ,  $-CH_2NH_2$  (d)  $-CH_2OCH_3$ ,  $-CN$ ,  $-C\equiv CH$ ,  $-CH_2CH_3$ 6.11 (a) *Z* (b) *E* (c) *Z* (d) *E*6.12 (a) 2-Methylpropene (b) *E* isomer (c) 1-Methylcyclohexene

6.14 (a) Chlorocyclohexane (b) 2-Bromo-2-methylpentane (c) 2-Iodopentane

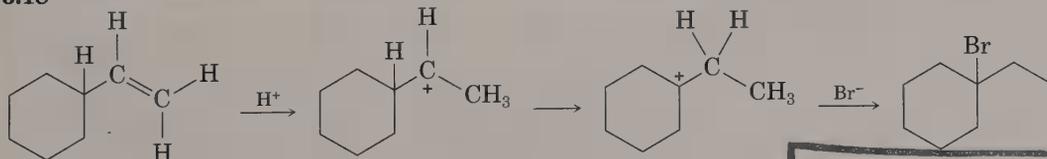
(d) 1-Bromo-1-methylcyclohexane

6.15 (a) Cyclopentene (b) 3-Hexene (c) 1-Ethylcyclohexene or ethylidenecyclohexane



6.17 The second step is exothermic; transition state resembles the carbocation.

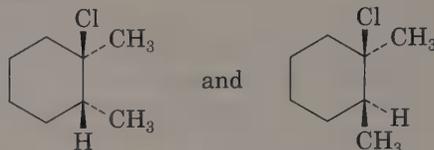
6.18



## CHAPTER 7

7.1 2-Methyl-2-butene and 2-methyl-1-butene  
 7.3 *trans*-1,2-Dichloro-1,2-dimethylcyclohexane

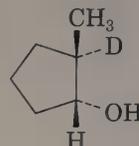
7.2 Five  
 7.4



7.5 *trans*-2-Bromocyclopentanol 7.6 Markovnikov 7.7  $\delta^+ \text{I} - \text{N}_3^{\delta-}$

7.8 (a) 2-Pentanol (b) 2-Methyl-2-pentanol

7.9 (a) 2-Methyl-1-hexene or 2-methyl-2-hexene (b) Cyclohexylethylene 7.10



7.11 (a) 3-Methyl-1-butene (b) 2-Methyl-2-butene (c) Methylene cyclohexane

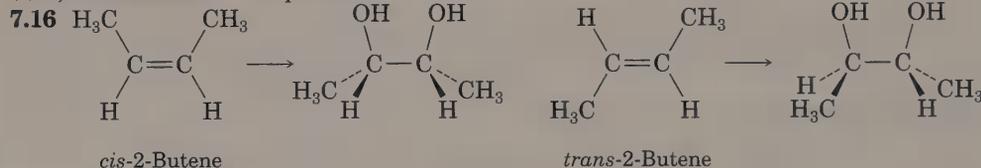
7.12 (a) 2-Methyl-1-butene + HBr (b) 1-Butene + HBr + peroxides

(c) 3-Methyl-3-hexene + HBr + peroxides

7.14 (a) 2-Methylpentane (b) 1,1-Dimethylcyclopentane

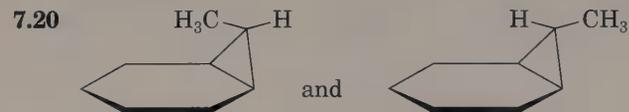
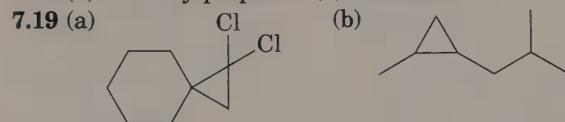
7.15 (a) 1-Methylcyclohexene + OsO<sub>4</sub> (b) 2-Methyl-2-pentene + OsO<sub>4</sub>

(c) 1,3-Butadiene + OsO<sub>4</sub>



7.17 (a) CH<sub>3</sub>COCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>COOH (b) CH<sub>3</sub>COCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CHO

7.18 (a) 2-Methylpropene (b) 3-Hexene



## CHAPTER 8

8.1 (a) 2,5-Dimethyl-3-hexyne (b) 3,3-Dimethyl-1-butyne (c) 2,4-Octadiene-6-yne  
 (d) 3,3-Dimethyl-4-octyne (e) 2,5,5-Trimethyl-3-heptyne (f) 6-Isopropylcyclodecyne

8.2 1-Hexyne, 2-hexyne, 3-hexyne, 3-methyl-1-pentyne, 4-methyl-1-pentyne, 4-methyl-2-pentyne, 3,3-dimethyl-1-butyne

8.3 (a) 1,1,2,2-Tetrachloropentane (b) 1-Bromo-1-cyclopentylethylene (c) 2-Bromo-2-heptene and 3-bromo-2-heptene 8.4 4-Octanone, 2-methyl-4-octanone, and 7-methyl-4-octanone

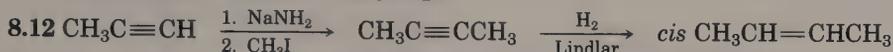
8.5 (a) 1-Pentyne (b) 3-Hexyne 8.6 (a) Phenylacetylene (b) 2,5-Dimethyl-3-hexyne

8.7 2-Methyl-2-butene 8.8 2-Methyl-2-butanol

8.9 (a) 2-Octyne + Li/NH<sub>3</sub> (b) 3-Heptyne + H<sub>2</sub>/Lindlar

(c) 3-Methyl-1-pentyne + H<sub>2</sub>/Lindlar 8.10 (b)

- 8.11 (a) 1-Pentyne +  $\text{CH}_3\text{I}$  or propyne +  $\text{CH}_3\text{CH}_2\text{CH}_2\text{I}$  (b) 3-Methyl-1-butyne +  $\text{CH}_3\text{CH}_2\text{I}$   
 (c) Cyclohexylacetylene +  $\text{CH}_3\text{I}$  (d) 4-Methyl-1-pentyne +  $\text{CH}_3\text{I}$  or propyne +  $(\text{CH}_3)_2\text{CHCH}_2\text{I}$   
 (e) 3-3-Dimethyl-1-butyne +  $\text{CH}_3\text{CH}_2\text{I}$



- 8.13 (a) Phenylacetylene (b)  $\text{CH}_3(\text{CH}_2)_7\text{C}\equiv\text{C}(\text{CH}_2)_7\text{C}\equiv\text{C}(\text{CH}_2)_7\text{CH}_3$

- 8.14 (a)  $\text{KMnO}_4, \text{H}_2\text{O}$  (b)  $\text{H}_2/\text{Lindlar}$  (c) 1.  $\text{H}_2/\text{Lindlar}$ ; 2.  $\text{HBr}$

- (d) 1.  $\text{H}_2/\text{Lindlar}$ ; 2.  $\text{BH}_3$ ; 3.  $\text{NaOH}, \text{H}_2\text{O}_2$  (e) 1.  $\text{H}_2/\text{Lindlar}$ ; 2.  $\text{Cl}_2$

- 8.15 (a) 1.  $\text{NaNH}_2$ ; 2.  $\text{CH}_3(\text{CH}_2)_7\text{C}\equiv\text{CH}$ ; 3. 2  $\text{H}_2/\text{Pd}$

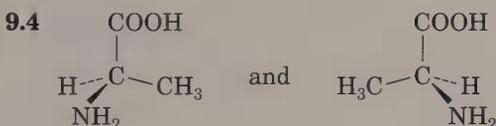
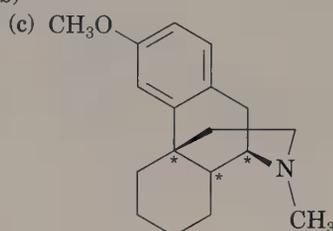
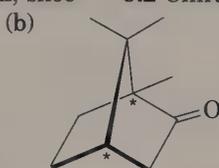
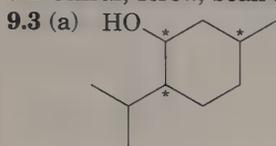
- (b) 1.  $\text{NaNH}_2$ ; 2.  $(\text{CH}_3)_3\text{CCH}_2\text{CH}_2\text{I}$ ; 3. 2  $\text{H}_2/\text{Pd}$

- (c) 1.  $\text{NaNH}_2$ ; 2.  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{I}$ ; 3.  $\text{HBR}_2$ ; 4.  $\text{H}_2\text{O}_2, \text{NaOH}$

- (d) 1.  $\text{NaNH}_2$ ; 2.  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{I}$ ; 3.  $\text{HgSO}_4, \text{H}_3\text{O}^+$

## CHAPTER 9

- 9.1 Chiral; screw, bean stalk, shoe 9.2 Chiral; (b)



- 9.5 Direction can be found by varying the concentration. 9.6  $+16^\circ$

- 9.7 (a)  $-\text{Br}$ ,  $-\text{CH}_2\text{CH}_2\text{OH}$ ,  $-\text{CH}_2\text{CH}_3$ ,  $-\text{H}$  (b)  $-\text{OH}$ ,  $-\text{CO}_2\text{CH}_3$ ,  $-\text{CO}_2\text{H}$ ,  $-\text{CH}_2\text{OH}$

- (c)  $-\text{NH}_2$ ,  $-\text{CN}$ ,  $-\text{CH}_2\text{NHCH}_3$ ,  $-\text{CH}_2\text{NH}_2$  (d)  $-\text{Br}$ ,  $-\text{Cl}$ ,  $-\text{CH}_2\text{Br}$ ,  $-\text{CH}_2\text{Cl}$

- 9.8 (a) *S* (b) *S* (c) *R*

- 9.9 (a) *R,R* (b) *S,R* (c) *R,S* (d) *S,S*; (a) and (d) are enantiomers and are diastereomeric to (b) and (c). 9.10 *R,R* 9.11 Meso: (a) and (d) 9.12 Meso: (a) and (c)

- 9.13 Five stereogenic centers; 32 stereoisomers

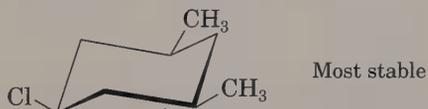
- 9.14 A and B are identical; C and D are identical. 9.15 (a) Enantiomers (b) Enantiomers

- 9.16 (a) *S* (b) *S* (c) *R* 9.17 Racemic mixture 9.19 Racemic mixture

- 9.20 Non-50:50 mixture of two racemic pairs; optically inactive

- 9.21 Non-50:50 mixture of two racemic pairs; optically inactive

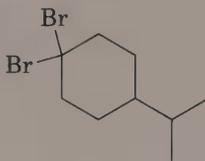
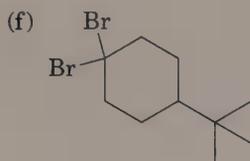
- 9.22 Four stereoisomers



- 9.23 Three stereoisomers: one pair of enantiomers and one meso compound

## CHAPTER 10

- 10.1 (a) 1-Iodobutane (b) 1-Chloro-3-methylbutane (c) 1,5-Dibromo-2,2-dimethylpentane  
 (d) 1,3-Dichloro-3-methylbutane (e) 1-Chloro-3-ethyl-4-iodopentane  
 (f) 2-Bromo-5-chlorohexane

10.2 (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{C}(\text{CH}_3)_2\text{CHClCH}_3$  (b)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CCl}_2\text{CH}(\text{CH}_3)_2$ (c)  $\text{CH}_3\text{CH}_2\text{CBr}(\text{CH}_2\text{CH}_3)_2$  (d)(e)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{CH}(\text{Cl})\text{CH}_3$ 

10.3 1-Chloro-2-methylpentane, 2-chloro-2-methylpentane, 3-chloro-2-methylpentane, 2-chloro-4-methylpentane, 1-chloro-4-methylpentane; the first, third, and fourth compounds are chiral.

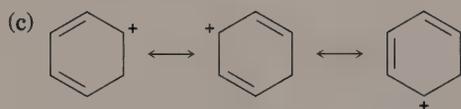
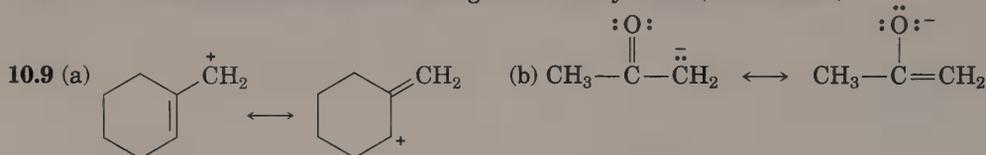
10.4 1-Chloro-2-methylbutane (29%), 1-chloro-3-methylbutane (14%), 2-chloro-2-methylbutane (24%), 2-chloro-3-methylbutane (33%)

10.5 For Cl;  $\Delta H^\circ = -8$  kcal/mol; for Br $\cdot$ ,  $\Delta H^\circ = +7$  kcal/mol. Bromination is more selective.

10.6 The intermediate allylic radical reacts at the more accessible site.

10.7 (a) 3-Bromo-5-methylcycloheptene and 3-bromo-6-methylcycloheptene (b) Four products

10.8 A C-H bond next to an aromatic ring is unusually weak (85 kcal/mol).

10.11 (a) 2-Methyl-2-propanol + HCl (b) 4-Methyl-2-pentanol + PBr<sub>3</sub>(c) 5-Methyl-1-pentanol + PBr<sub>3</sub> (d) 2,4-Dimethyl-2-hexanol + HCl10.12 Both reactions occur. 10.13 React Grignard reagent with D<sub>2</sub>O.

10.14 The -OH proton reacts with the Grignard reagent.

10.15 (a) 1. NBS; 2. (CH<sub>3</sub>)<sub>2</sub>CuLi (b) 1. Li; 2. CuI; 3. CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Br(c) 1. HBr, peroxides; 2. Li; 3. CuI; 4. CH<sub>3</sub>(CH<sub>2</sub>)<sub>3</sub>CH<sub>2</sub>Br

## CHAPTER 11

11.1 (*R*)-1-Methylpentyl acetate,  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)\text{OOCCH}_3$  11.2 (*S*)-2-Butanol

11.3 Back-side attack is too hindered.

11.4 (a) 1-Iodobutane (b) 1-Butanol (c) 1-Hexyne (d) Butylamine

11.5 The quinuclidine nitrogen is exposed because the substituent groups are tied back.

11.6 (a) (CH<sub>3</sub>)<sub>2</sub>N<sup>-</sup> (b) (CH<sub>3</sub>)<sub>3</sub>N (c) H<sub>2</sub>S11.7 CH<sub>3</sub>OTos > CH<sub>3</sub>Br > (CH<sub>3</sub>)<sub>2</sub>CHCl > (CH<sub>3</sub>)<sub>3</sub>CCl

11.8 Intermediate between protic and polar solvents

11.9 Racemic 1-ethyl-1-methyloctyl acetate 11.10 90.1% racemization, 9.9% inversion

11.11 H<sub>2</sub>C=CHCH(Br)CH<sub>3</sub> > CH<sub>3</sub>CH(Br)CH<sub>3</sub> > CH<sub>3</sub>CH<sub>2</sub>Br > H<sub>2</sub>C=CHBr

11.12 A common intermediate is involved.

- 11.13 The rate-limiting step does not involve the nucleophile.  
 11.14 The carbon atom is constrained from being planar.  
 11.15 (a) 2-Methyl-2-pentene (b) 2,3,5-Trimethyl-2-hexene (c) Ethylidenecyclohexane  
 11.16 (Z)-1-Bromo-1,2-diphenylethylene  
 11.17 Transition state leading to trans product is less hindered.  
 11.18 Cis isomer reacts faster because the bromine is axial. 11.19 S<sub>N</sub>2 (b) E2 (c) S<sub>N</sub>1

## CHAPTER 12

- 12.1 (a) C<sub>6</sub>H<sub>14</sub>, C<sub>5</sub>H<sub>10</sub>O, C<sub>4</sub>H<sub>6</sub>O<sub>2</sub>, C<sub>3</sub>H<sub>2</sub>O<sub>3</sub> (b) C<sub>9</sub>H<sub>20</sub>, C<sub>8</sub>H<sub>16</sub>O, C<sub>7</sub>H<sub>12</sub>O<sub>2</sub>, C<sub>6</sub>H<sub>8</sub>O<sub>3</sub>, C<sub>5</sub>H<sub>4</sub>O<sub>4</sub>  
 (c) C<sub>11</sub>H<sub>24</sub>, C<sub>10</sub>H<sub>20</sub>O, C<sub>9</sub>H<sub>16</sub>O<sub>2</sub>, C<sub>8</sub>H<sub>12</sub>O<sub>3</sub>, C<sub>7</sub>H<sub>8</sub>O<sub>4</sub>, C<sub>6</sub>H<sub>4</sub>O<sub>5</sub>  
 12.2 C<sub>15</sub>H<sub>22</sub>O, C<sub>14</sub>H<sub>18</sub>O<sub>2</sub>, C<sub>13</sub>H<sub>14</sub>O<sub>3</sub>, C<sub>12</sub>H<sub>10</sub>O<sub>4</sub>, C<sub>11</sub>H<sub>6</sub>O<sub>5</sub> 12.3 100% and 6.72%  
 12.4 An odd number of nitrogens implies an odd number of hydrogens. 12.5 C<sub>5</sub>H<sub>5</sub>N  
 12.6 C<sub>4</sub>H<sub>9</sub>, *tert*-butyl cation 12.7 (a) 2-Methyl-2-pentene (b) 2-Hexene  
 12.8 X-ray energy is higher. 12.9  $\lambda = 10^{-4}$  cm is higher in energy.  
 12.10 (a)  $5.7 \times 10^5$  kcal/mol (b)  $9.5 \times 10^3$  kcal/mol (c)  $5.7 \times 10^2$  kcal/mol (d) 67 kcal/mol  
 (e) 1.4 kcal/mol (f)  $9.5 \times 10^{-3}$  kcal/mol (g)  $9.76 \times 10^{-6}$  kcal/mol  
 12.11 (a) 3225 cm<sup>-1</sup> (b) 1710 cm<sup>-1</sup> (c) 4.44  $\mu$ m (d) 10.3  $\mu$ m  
 12.12 (a) Ketone or aldehyde (b) Nitro (c) Carboxylic acid  
 12.13 (a) CH<sub>3</sub>CH<sub>2</sub>OH has an -OH absorption. (b) 1-Hexene has a double-bond absorption.  
 (c) CH<sub>3</sub>CH<sub>2</sub>COOH has a very broad -OH absorption.  
 12.15 (a) 1715 cm<sup>-1</sup> (b) 3300, 2100, 1725 cm<sup>-1</sup>

## CHAPTER 13

- 13.1  $5.3 \times 10^{-6}$  kcal/mol for <sup>19</sup>F versus  $5.7 \times 10^{-6}$  kcal/mol for <sup>1</sup>H  
 13.2  $9.5 \times 10^{-6}$  kcal/mol  
 13.3 (a) <sup>1</sup>H one; <sup>13</sup>C two (b) <sup>1</sup>H one; <sup>13</sup>C one (c) <sup>1</sup>H one; <sup>13</sup>C one (d) <sup>1</sup>H two; <sup>13</sup>C four  
 (e) <sup>1</sup>H two; <sup>13</sup>C three (f) <sup>1</sup>H two; <sup>13</sup>C three  
 13.4 The vinylic C-H protons are nonequivalent. 13.5 (a) 126 Hz (b) 2.1  $\delta$  (c) 210 Hz  
 13.6 (a) 7.27  $\delta$  (b) 3.05  $\delta$  (c) 3.47  $\delta$  (d) 5.30  $\delta$   
 13.8 (a) Four (b) Seven (c) Four (d) Five  
 13.9 (a) 1,3-Dimethylcyclopentene (b) 2-Methylpentane (c) 1-Chloro-2-methylpropane  
 13.10 Three lines; 1:1:2  
 13.11 (a) Three (b) -CH<sub>3</sub> carbon has four lines; -CH= carbon has two lines; =CH<sub>2</sub> carbon has three lines. 13.12 (a) Three (b) Integration shows 1:2:2 ratio.  
 13.13 1-Bromo-2-methylpropane has three kinds of carbons; 2-bromo-2-methylpropane has two kinds of carbons. 13.14 (a) Two (b) Four (c) Three (d) Four (e) Five (f) Three  
 13.15 (a) 1.4  $\delta$  (b) 2.17  $\delta$  (c) 7.37  $\delta$  (d) 9.70  $\delta$  (e) 5.30  $\delta$  (f) 2.12  $\delta$   
 13.17 (a) CHBr<sub>2</sub><sup>-</sup>, quartet; -CH<sub>3</sub>, doublet (b) CH<sub>3</sub>O, singlet; -OCH<sub>2</sub><sup>-</sup>, triplet; -CH<sub>2</sub>Br, triplet (c) ClCH<sub>2</sub><sup>-</sup>, triplet; -CH<sub>2</sub><sup>-</sup>, quintet  
 (d) CH<sub>3</sub><sup>-</sup>, triplet; -CH<sub>2</sub><sup>-</sup>, quartet; -CH<sup>-</sup>, septet; (CH<sub>3</sub>)<sub>2</sub>, doublet  
 13.18 (a) CH<sub>3</sub>OCH<sub>3</sub> (b) CH<sub>3</sub>CHClCH<sub>3</sub> (c) ClCH<sub>2</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>2</sub>Cl (d) CH<sub>3</sub>CH<sub>2</sub>COOCH<sub>3</sub>  
 13.19 CH<sub>3</sub>CH<sub>2</sub>OCH<sub>2</sub>CH<sub>3</sub>

## CHAPTER 14

- 14.1 Conjugated: (b), (c), (d), (f)  
 14.2 Expected  $\Delta H^\circ_{\text{hydrog}}$  for allene—60.6 kcal/mol. Allene is less stable than either a conjugated or a nonconjugated diene.

- 14.3 1-Chloro-2-pentene, 3-chloro-1-pentene, 4-chloro-1-pentene, 4-chloro-2-pentene  
 14.4 4-Chloro-2-pentene predominates in both.  
 14.5 The intermediate carbocation is more stable.  
 14.7 Interconversion occurs by  $S_N1$  dissociation to a common intermediate cation.  
 14.8 The double bond is more highly substituted. 14.9 Good dienophiles: (a), (d)  
 14.10 *s*-cis: (a). Compound (c) can rotate to *s*-cis.  
 14.11 Increasing the ring size moves the terminal diene carbons farther apart.  
 14.12 72–140 kcal/mol 14.13 UV energy is higher than IR or NMR energy.  
 14.14  $1.47 \times 10^{-5}$  M 14.15 UV absorptions: all except (a)

## CHAPTER 15

- 15.2 (a) Meta (b) Para (c) Ortho  
 15.3 (a) *m*-Bromochlorobenzene (b) (3-Methylbutyl)benzene or isobutylbenzene  
 (c) *p*-Bromoaniline (d) 2,4-Dichlorotoluene (e) 1-Ethyl-2,4-dinitrobenzene  
 (f) 1,2,3,5-Tetramethylbenzene  
 15.5 (a) 1-Bromo-2-chlorobenzene or *o*-bromochlorobenzene (b) 2,4-Dinitrotoluene  
 (c) 1-Bromo-4-methylbenzene or *p*-bromotoluene (d) 2-Chloro-1,4-dimethylbenzene  
 15.6 Four: 1,2,3-tribromobenzene, 1,3,5-tribromobenzene, and two 1,2,4-tribromobenzenes  
 15.7 One monobromo and four dibromo derivatives for Ladenburg benzene; three monobromo  
 and six dibromo derivatives for Dewar benzene.  
 15.8 The 3:2:1 experimental ratio corresponds to the theoretical ratio.  
 15.10 Cyclodecapentaene is not flat.  
 15.11 Five resonance structures; all C–C bonds are equivalent; one resonance line in both  $^1\text{H}$   
 and  $^{13}\text{C}$  NMR spectra.  
 15.12 The cyclooctatetraenyl dianion is aromatic (eight pi electrons) and flat.  
 15.13 The singly bonded nitrogen contributes two pi electrons and the doubly bonded  
 nitrogen contributes one pi electron to the aromatic sextet.  
 15.14 The oxygen contributes two pi electrons to the aromatic sextet.  
 15.16 Two resonance structures have a single bond between C2 and C3; only one resonance  
 structure has a single bond between C1 and C2. 15.17 Five

## CHAPTER 16

- 16.1 *o*-, *m*-, and *p*-Bromotoluene 16.2  $\delta^+\text{I}-\text{Cl}^{\delta-}$   
 16.4  $\text{H}^+$  is the electrophile; the mechanism of desulfonation is the exact reverse of  
 sulfonation. 16.5 *o*-Xylene has two kinds of positions available, but *p*-xylene has only one.  
 16.6 Three 16.7  $\text{D}^+$  does electrophilic substitutions on the ring. 16.8 *tert*-Butylbenzene  
 16.9 Protonation of 2-methylpropene yields the *tert*-butyl cation.  
 16.10 No rearrangement: (a), (b), (e)  
 16.11 (a) Phenol > toluene > benzene > nitrobenzene  
 (b) Phenol > benzene > chlorobenzene > benzoic acid  
 (c) Aniline > benzene > bromobenzene > benzaldehyde  
 16.12 Alkylbenzenes are more reactive than benzene itself, but acylbenzenes are less  
 reactive. 16.13 Bromination is most favored at the ortho and para positions.  
 16.14 Nitration is most favored at the meta position.  
 16.15 (a) *o*- and *p*-Bromoaniline (b) *m*-Bromonitrobenzene (c) *o*- and *p*-Chlorophenol  
 (d), (e) *o*- and *p*-Chlorophenol  
 16.16 Toluene; the trifluoromethyl group is electron-withdrawing.  
 16.17 The nitrogen electrons are donated to the nearby carbonyl group.  
 16.18 Phenoxide ion > phenol > phenyl acetate 16.19 Meta  
 16.21 (a) Ortho and para to  $-\text{OCH}_3$  (b) Ortho and para to  $-\text{NH}_2$  (c) Ortho and para to  $-\text{Cl}$

- 16.23 Only one benzyne intermediate can form from *p*-bromotoluene; two different benzyne intermediates can form from *m*-bromotoluene.
- 16.24 (a) 1,2-Benzenedicarboxylic acid (b) *m*-Nitrobenzoic acid
- 16.25 1.  $\text{CH}_3\text{CH}_2\text{Cl}$ ,  $\text{AlCl}_3$ ; 2. NBS; 3. KOH, ethanol
- 16.26 A benzyl radical is more stable than a primary alkyl radical by 13 kcal/mol and is similar in stability to an allyl radical.
- 16.27 1,3-Dimethyl-1,4-cyclohexadiene and 1,5-dimethyl-1,4-cyclohexadiene
- 16.28 1.  $\text{PhCOCl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{H}_2/\text{Pd}$
- 16.29 (a) 1.  $\text{HNO}_3$ ,  $\text{H}_2/\text{Pd}$ ; 2.  $\text{Cl}_2$ ,  $\text{FeCl}_3$  (b) 1.  $\text{CH}_3\text{CH}_2\text{COCl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{Cl}_2$ ,  $\text{FeCl}_3$ ; 3.  $\text{H}_2/\text{Pd}$   
(c) 1.  $\text{Cl}_2$ ,  $\text{FeCl}_3$ ; 2.  $\text{CH}_3\text{CH}_2\text{COCl}$ ,  $\text{AlCl}_3$ ; 3.  $\text{H}_2/\text{Pd}$
- 16.30 (a) Friedel–Crafts acylation does not occur on a deactivated ring.  
(b) Rearrangement occurs during Friedel–Crafts alkylation with primary halides; chlorination occurs ortho to the alkyl group.

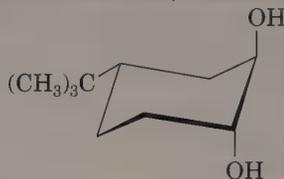
## CHAPTER 17

- 17.1 (a) 5-Methyl-2,4-hexanediol (b) 2-Methyl-4-phenyl-2-butanol  
(c) 4,4-Dimethylcyclohexanol (d) *trans*-2-Bromocyclopentanol
- 17.3 Hydrogen bonding is more difficult in hindered alcohols. 17.4 *p*-Nitrobenzyl alcohol
- 17.5 (a) 2-Methyl-4-phenyl-1-butanol (b) 2-Methyl-2-pentanol (c) *meso*-5,6-Decanediol
- 17.6 Cyclohexane < cyclohexanone < 1-methoxycyclohexene < benzene
- 17.7 (a) Neither (b) Oxidation (c) Reduction (d) Neither
- 17.8 (a)  $\text{NaBH}_4$  (b)  $\text{LiAlH}_4$  (c)  $\text{LiAlH}_4$  (d)  $\text{H}_2/\text{Pd}$
- 17.9 (a) Benzaldehyde or benzoic acid (or ester) (b) Acetophenone (c) Cyclohexanone  
(d) 2-Methylpropanal or 2-methylpropanoic acid (or ester)
- 17.10 (a) 1-Methylcyclopentanol (b) 1,1-Diphenylethanol (c) 3-Methyl-3-hexanol
- 17.11 (a) Acetone +  $\text{CH}_3\text{MgBr}$  or ethyl acetate + 2  $\text{CH}_3\text{MgBr}$  (b) Cyclohexanone +  $\text{CH}_3\text{MgBr}$   
(c) 3-Pentanone +  $\text{CH}_3\text{MgBr}$  or 2-butanone +  $\text{CH}_3\text{CH}_2\text{MgBr}$  or ethyl acetate + 2  $\text{CH}_3\text{CH}_2\text{MgBr}$   
(d) 2-Butanone +  $\text{PhMgBr}$  or ethyl phenyl ketone +  $\text{CH}_3\text{MgBr}$  or acetophenone + 2  $\text{CH}_3\text{CH}_2\text{MgBr}$   
(e) Formaldehyde +  $\text{PhMgBr}$  17.12 The  $-\text{OH}$  proton reacts with the Grignard reagent.
- 17.13 (a) 2-Methyl-2-pentene (b) 3-Methylcyclohexene (c) 1-Methylcyclohexene
- 17.15 (a) 1-Phenylethanol (b) 2-Methyl-1-propanol (c) Cyclopentanol
- 17.16 (a) Hexanoic acid, hexanal (b) 2-Hexanone (c) Hexanoic acid, no reaction
- 17.17  $\text{S}_{\text{N}}2$  reaction of  $\text{F}^-$  on silicon
- 17.18 Disappearance of  $-\text{OH}$  absorption; appearance of  $\text{C}=\text{O}$
- 17.19 (a) Singlet (b) Doublet (c) Triplet (d) Doublet (e) Doublet (f) Singlet
- 17.20 (a) 2-Butanethiol (b) 2,2,6-Trimethyl-4-heptanethiol (c) 2-Cyclopentene-1-thiol
- 17.21 (a) 1.  $\text{LiAlH}_4$ ; 2.  $\text{PBr}_3$ ; 3.  $(\text{H}_2\text{N})_2\text{C}=\text{S}$ ; 4.  $\text{H}_2\text{O}$ ,  $\text{NaOH}$   
(b) 1.  $\text{HBr}$ ; 2.  $(\text{H}_2\text{N})_2\text{C}=\text{S}$ ; 3.  $\text{H}_2\text{O}$ ,  $\text{NaOH}$

## CHAPTER 18

- 18.1 (a) Diisopropyl ether (b) Cyclopentyl propyl ether  
(c) *p*-Bromoanisole or 4-bromo-1-methoxybenzene (d) 1-Methoxycyclohexene  
(e) Ethyl isobutyl ether (f) Allyl vinyl ether
- 18.2 Mixtures of products results: diethyl ether, dipropyl ether, ethyl propyl ether.
- 18.4 (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{O}^- + \text{CH}_3\text{I}$  (b)  $\text{PhO}^- + \text{CH}_3\text{I}$  (c)  $(\text{CH}_3)_2\text{CHO}^- + \text{PhCH}_2\text{Br}$   
(d)  $(\text{CH}_3)_3\text{CCH}_2\text{O}^- + \text{CH}_3\text{CH}_2\text{Br}$
- 18.5 (a) Bromoethane > 2-bromopropane > bromobenzene  
(b) Bromoethane > chloroethane > 1-iodopropene
- 18.7 (a) Either method (b) Alkoxymercuration method (c) Alkoxymercuration method  
(d) Williamson method 18.8  $\text{E1}$  reaction

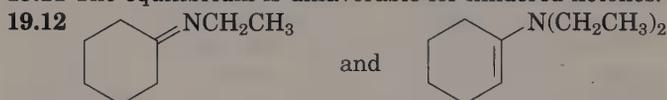
- 18.9  $\text{Br}^-$  and  $\text{I}^-$  are better nucleophiles than  $\text{Cl}^-$  is. 18.10 *meso*-2,3-Epoxybutane  
 18.11 Racemic 2,3-epoxybutane 18.12 Racemic 5,6-decanediol  
 18.13 *meso*-5,6-Decanediol 18.14



- 18.15 (a)  $\text{CH}_3\text{CH}_2\text{CH}(\text{OH})\text{CH}_2\text{NH}_2$  (b)  $\text{CH}_3\text{CH}_2\text{CH}^{18}\text{OHCH}_2\text{OH}$  18.17 1,2-Epoxybutane  
 18.18 (a) Ethyl methyl sulfide (b) *tert*-Butyl ethyl sulfide (c) *o*-(Dimethylthio)benzene  
 (d) *p*-(Phenylthio)toluene 18.19 Dimethyl sulfoxide is highly polar.

## CHAPTER 19

- 19.1 (a) 2-Methyl-3-pentanone (b) 3-Phenylpropanal (c) 2,6-Octanedione  
 (d) *trans*-2-Methylcyclohexancarbaldehyde (e) Pentanedial  
 (f) *cis*-2,5-Dimethylcyclohexanone (g) 4-Methyl-3-propyl-2-hexanone (h) 4-Hexenal  
 19.3 (a) 1. PCC (b) 1.  $\text{O}_3$ ; 2. Zn (c) 1. DIBAL  
 19.4 (a) 1.  $\text{Hg}(\text{OAc})_2$ ,  $\text{H}_3\text{O}^+$  (b) 1.  $\text{CH}_3\text{COCl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{Br}_2$ ,  $\text{FeBr}_3$   
 (c) 1. Mg; 2.  $\text{CH}_3\text{CHO}$ ; 3.  $\text{H}_3\text{O}^+$ ; 4. PCC (d) 1.  $\text{BH}_3$ ; 2.  $\text{H}_2\text{O}_2$ , NaOH; 3. PCC  
 19.5  $(\text{CH}_3)_2\text{C}(\text{OH})\text{CN}$  19.6 Carbonyl carbon in benzaldehyde is less electrophilic.  
 19.7 *p*-Nitrobenzaldehyde 19.8  $\text{CCl}_3\text{CH}(\text{OH})_2$   
 19.9 Labeled water adds reversibly to the carbonyl group.  
 19.10 The  $\text{S}_{\text{N}}2$  substitution product is an aldehyde hydrate that loses water.  
 19.11 The equilibrium is unfavorable for hindered ketones.



- 19.13 The steps are the exact reverse of the forward reaction.  
 19.15 (a) 1 equiv  $\text{NaBH}_4$  (b) 1. 1 equiv  $\text{HOCH}_2\text{CH}_2\text{OH}$ ,  $\text{H}^+$ ; 2.  $\text{NaBH}_4$ ; 3.  $\text{H}_3\text{O}^+$   
 19.16 (a) Wolff-Kishner reaction, (2) Clemmensen reduction, (3) Raney Ni reduction of dithioacetal 19.17 1. 1 equiv  $\text{HSCH}_2\text{CH}_2\text{SH}$ ; 2. Raney Ni  
 19.18 (a) Cyclohexanone +  $(\text{Ph})_3\text{P}=\text{CHCH}_3$  (b) 2-Cyclohexenone +  $(\text{Ph})_3\text{P}=\text{CH}_2$   
 (c) Acetone +  $(\text{Ph})_3\text{P}=\text{CHCH}_2\text{CH}_2\text{CH}_3$  (d) Acetone +  $(\text{Ph})_3\text{P}=\text{CHPh}$   
 (e) Benzaldehyde +  $(\text{Ph})_3\text{P}=\text{CHPh}$   
 19.19 Using trimethylphosphine would lead to a mixture of products.  
 19.20  $(\text{Ph})_3\text{P}=\text{CHC}(\text{CH}_3)=\text{CHCH}=\text{CHCH}=\text{C}(\text{CH}_3)\text{CH}=\text{P}(\text{Ph})_3$   
 19.21 Intramolecular Cannizzaro reaction  
 19.22 (a) 3-Buten-2-one +  $(\text{CH}_3\text{CH}_2\text{CH}_2)_2\text{CuLi}$  (b) 3-Methyl-2-cyclohexenone +  $(\text{CH}_3)_2\text{CuLi}$   
 (c) 4-*tert*-Butyl-2-cyclohexenone +  $(\text{CH}_3\text{CH}_2)_2\text{CuLi}$   
 19.24 (a)  $1715\text{ cm}^{-1}$  (b)  $1685\text{ cm}^{-1}$  (c)  $1750\text{ cm}^{-1}$  (d)  $1705\text{ cm}^{-1}$  (e)  $1715\text{ cm}^{-1}$   
 (f)  $1705\text{ cm}^{-1}$  19.25 3-Phenyl-3-cyclohexenone

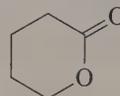
## CHAPTER 20

- 20.1 (a) 3-Methylbutanoic acid (b) 4-Bromopentanoic acid (c) 2,4-Hexadienoic acid  
 (d) 2-Ethylpentanoic acid (e) *cis*-1,3-Cyclopentanedicarboxylic acid  
 (f) 2-Phenylpropanoic acid 20.3 +22 kcal/mol, +6.42 kcal/mol  
 20.4 Dissolve the mixture in ether, then extract with aqueous NaOH, then acidify the aqueous layer and extract with ether. 20.5 52%

- 20.6 (a)  $\text{CH}_3\text{CH}_2\text{COOH} < \text{BrCH}_2\text{COOH} < \text{FCH}_2\text{COOH}$   
 (b) *p*-Methoxybenzoic acid < benzoic acid < *p*-nitrobenzoic acid  
 (c)  $\text{CH}_3\text{CH}_2\text{NH}_2 < \text{CH}_3\text{CH}_2\text{OH} < \text{CH}_3\text{CH}_2\text{COOH}$   
 20.7 The dianion is destabilized by repulsion between charges.  
 20.8 The dianions become more stable as charges move farther apart.      20.9 More reactive  
 20.10 (a) *p*-Methylbenzoic acid < benzoic acid < *p*-chlorobenzoic acid  
 (b) Acetic acid < benzoic acid < *p*-nitrobenzoic acid  
 20.12 (a) Grignard carboxylation    (b) Grignard carboxylation    (c) Either method

## CHAPTER 21

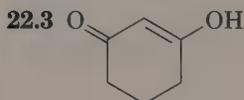
- 21.1 (a) 4-Methylpentanoyl chloride    (b) Cyclohexylacetamide    (c) 2-Methylbutanenitrile  
 (d) Benzoic anhydride    (e) Isopropyl cyclopentanecarboxylate  
 (f) Cyclopentyl 2-methylpropanoate    (g) 4-Pentenamide    (h) 2-Ethylbutanenitrile  
 (i) 2,3-Dimethyl-2-butenoyl chloride    (j) Bis(trifluoroacetic) anhydride  
 21.2 (a) 2-Pentenitrile    (b) *N*-Methylbutanamide    (c) 2,4-Dimethylpentanoyl chloride  
 (d) Methyl 1-methylcyclohexanecarboxylate  
 21.3 (a) Acetyl chloride > methyl acetate > acetamide  
 (b) Hexafluoroisopropyl acetate > 2,2,2-trichloroethyl acetate > methyl acetate  
 21.4 The electron-withdrawing trifluoromethyl group polarizes the carbonyl carbon.  
 21.5 (a) Acetic acid + 1-butanol    (b) Butanoic acid + methanol      21.6



- 21.7  $\text{S}_{\text{N}}2$  reaction  
 21.8 (a) Propanoyl chloride + methanol    (b) Acetyl chloride + ethanol  
 (c) Benzoyl chloride + ethanol      21.9 Benzoyl chloride + cyclohexanol  
 21.11 Equilibrium favors reactants.  
 21.12 (a) Propanoyl chloride + methylamine    (b) Benzoyl chloride + dimethylamine  
 (c) Propanoyl chloride + ammonia  
 21.13 (a) Benzoyl chloride +  $[(\text{CH}_3)_2\text{CH}]_2\text{CuLi}$  or 2-methylpropanoyl chloride +  $\text{Ph}_2\text{CuLi}$   
 (b) Propenoyl chloride +  $(\text{CH}_3\text{CH}_2\text{CH}_2)_2\text{CuLi}$  or butanoyl chloride +  $(\text{H}_2\text{C}=\text{CH})_2\text{CuLi}$   
 21.14 Acid chlorides are more reactive than ketones toward Grignard reaction.  
 21.15 Monomethyl benzene-1,2-dicarboxylate  
 21.17 If no added base were present, half of the reactant amine would form a salt.  
 21.18 Reduce labeled acetic acid with  $\text{LiAlH}_4$  to give  $\text{CH}_3\text{CH}_2^{18}\text{OH}$ , and then make ester.  
 21.19 Reaction of an acid with an alkoxide ion gives the carboxylate salt.  
 21.20  $\text{HOCH}_2\text{CH}_2\text{CH}_2\text{CHO}$   
 21.21 (a)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}(\text{CH}_3)\text{CH}_2\text{OH}$     (b)  $\text{PhOH} + \text{PhCH}_2\text{OH}$   
 21.22 (a) Ethyl benzoate + 2  $\text{CH}_3\text{MgBr}$     (b) Ethyl acetate + 2  $\text{PhMgBr}$   
 (c) Ethyl butanoate + 2  $\text{CH}_3\text{CH}_2\text{MgBr}$   
 21.23 (a)  $\text{H}_2\text{O}$ ,  $\text{NaOH}$     (b) Benzoic acid +  $\text{BH}_3$     (c)  $\text{LiAlH}_4$   
 21.24 1.  $\text{Mg}$ ; 2.  $\text{CO}_2$ , then  $\text{H}_3\text{O}^+$ ; 3.  $\text{SOCl}_2$ ; 4.  $(\text{CH}_3)_2\text{NH}$ ; 5.  $\text{LiAlH}_4$   
 21.26 (a)  $\text{CH}_3\text{CH}_2\text{CN} + \text{CH}_3\text{CH}_2\text{MgBr}$   
 (b)  $\text{CH}_3\text{CH}_2\text{CN} + (\text{CH}_3)_2\text{CHMgBr}$  or  $(\text{CH}_3)_2\text{CHCN} + \text{CH}_3\text{CH}_2\text{MgBr}$   
 (c)  $(\text{CH}_3)_2\text{CHCN} + \text{DIBAL}$     (d)  $\text{PhCN} + \text{CH}_3\text{MgBr}$  or  $\text{CH}_3\text{CN} + \text{PhMgBr}$   
 (e) Cyclohexanecarbonitrile +  $\text{C}_6\text{H}_{11}\text{MgBr}$       21.27 1.  $\text{NaCN}$ ; 2.  $\text{CH}_3\text{CH}_2\text{MgBr}$   
 21.28 The imine anion is unreactive because of its negative charge.  
 21.29 (a) Ester    (b) Acid chloride    (c) Carboxylic acid    (d) Nitrile  
 (e) Aliphatic ketone or cyclohexanone  
 21.30 (a)  $\text{CH}_3\text{CH}_2\text{CN}$     (b)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{COOCH}_2\text{CH}_3$  or  $\text{CH}_3\text{COOCH}_2\text{CH}_2\text{CH}_2\text{CH}_3$   
 (c)  $\text{CH}_3\text{CON}(\text{CH}_3)_2$     (d)  $\text{CH}_3\text{CH}=\text{CHCOCl}$  or  $\text{H}_2\text{C}=\text{C}(\text{CH}_3)\text{COCl}$

## CHAPTER 22

22.2 (a) Four (b) Three (c) Three (d) Three (e) Four (f) Five (g) Three



Most stable

22.5 The enol intermediate is achiral.

22.6 No racemization 22.7 1. Br<sub>2</sub>; 2. Pyridine, Δ 22.8 Racemic22.10 (a) CH<sub>3</sub>CH<sub>2</sub>CHO (b) (CH<sub>3</sub>)<sub>3</sub>COCH<sub>3</sub> (c) CH<sub>3</sub>COOH (d) PhCONH<sub>2</sub>(e) CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>CN (f) CH<sub>3</sub>CO(CH<sub>3</sub>)<sub>2</sub> (g) -COCH<sub>2</sub>CO-

22.11 The enolate ion intermediate is achiral. 22.12 No

22.13 Enolate ion formation is rate-limiting.

22.14 HX is produced as a by-product of ketone halogenation.

22.17 Selenenylation is analogous to halogenation.

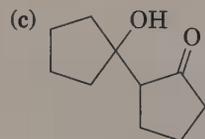
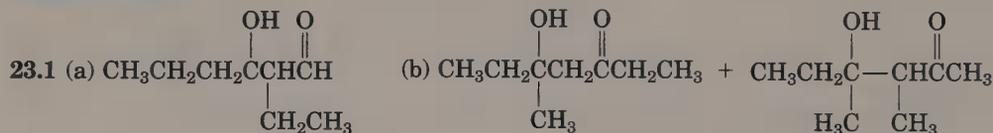
22.18 (a) CH<sub>3</sub>CH<sub>2</sub>Br (b) (CH<sub>3</sub>)<sub>2</sub>CHCH<sub>2</sub>CH<sub>2</sub>Br22.19 (a) Alkylate with PhCH<sub>2</sub>Br. (b) Alkylate with CH<sub>3</sub>CH<sub>2</sub>CH<sub>2</sub>Br.(c) Alkylate with (CH<sub>3</sub>)<sub>2</sub>CHCH<sub>2</sub>Br. (d) Alkylate with BrCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Br.

22.20 An alpha hydrogen is left after decarboxylation.

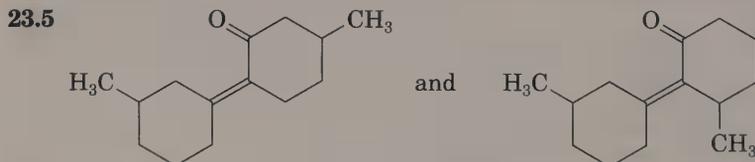
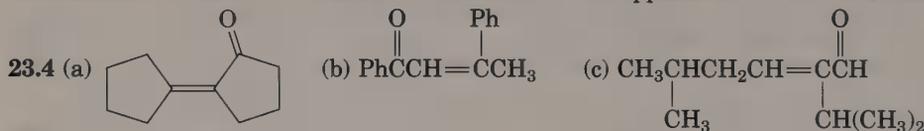
22.21 (a) (CH<sub>3</sub>)<sub>2</sub>CHCH<sub>2</sub>Br (b) PhCH<sub>2</sub>CH<sub>2</sub>Br 22.22 Alkylate with BrCH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>CH<sub>2</sub>Br.

22.23 Can't be prepared: (b), (c), (d)

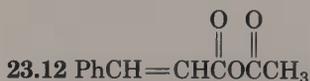
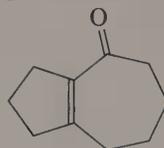
## CHAPTER 23

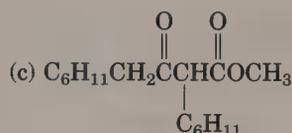
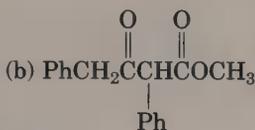
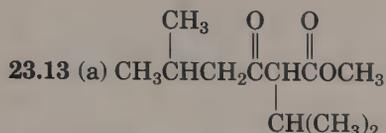


23.2 An enol 23.3 The reverse reaction is the exact opposite of the forward reaction.

23.6 Self-condensation products: (c) 23.7 1. NaOH; 2. LiAlH<sub>4</sub>; 3. H<sub>2</sub>/Pd 23.8 (a), (d)

23.9 The equilibrium is unfavorable because of ring strain. 23.10

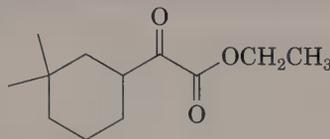




23.14 The cleavage reaction is the exact reverse of the forward reaction.

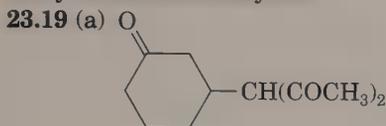
23.15 Yes;  $\text{EtO}_2\text{CCOCH}_2\text{CO}_2\text{Et}$

23.16

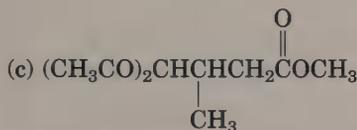


23.17 Ethyl 5-methyl-2-oxocyclohexanecarboxylate

23.18 Ethyl 4-methyl-2-oxocyclohexanecarboxylate and ethyl 6-methyl-2-oxocyclohexanecarboxylate



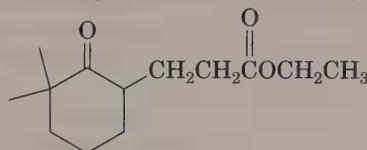
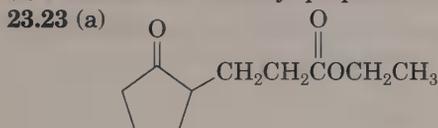
(b)  $(\text{CH}_3\text{CO})_2\text{CHCH}_2\text{CH}_2\text{CN}$



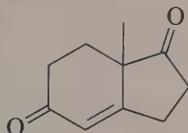
23.21 (a) Methyl acetoacetate + phenol vinyl ketone

(b) Methyl acetoacetate + 2-buten-3-one (c) Dimethyl malonate + propenenitrile

(d) Nitroethane + methyl propenoate (e) Dimethyl malonate + nitroethylene



23.25



## CHAPTER 24

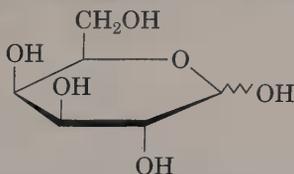
24.1 (a) Aldotetrose (b) Ketopentose (c) Ketohexose (d) Aldopentose

24.2 A, B, and C are the same 24.3 (a) *S* (b) *R* (c) *S*

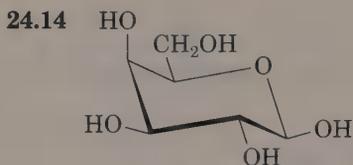
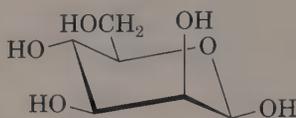
24.4 (a) *L*-erythrose;  $2S,3S$  (b) *D*-xylose;  $2R,3S,4R$  (c) *D*-xylulose;  $3S,4R$

24.5 *L*-Arabinose 24.7 There are 16 *D* and 16 *L* aldoseptoses.

24.9 24.12 36%  $\alpha$ -anomer; 64%  $\beta$ -anomer



24.13 28%  $\alpha$ -anomer; 72%  $\beta$ -anomer

 $\beta$ -D-Galactopyranose $\beta$ -D-Mannopyranose

24.17 D-Galactitol has a plane of symmetry.

24.19 D-Allaric acid has a symmetry plane, but D-glucaric acid does not.

24.20 D-Allose and D-galactose yield meso aldaric acids; the other six D-hexoses yield optically active aldaric acids.

24.21 D-Allose + D-altrose

24.22 L-Xylose

24.23 D-Xylose and D-lyxose

## CHAPTER 25

25.1 (a) Primary (b) Tertiary (c) Quaternary (d) Secondary (e) Secondary

25.2 (a) *N*-Methylethylamine (b) Tricyclohexylamine

(c) *N*-Methyl-*N*-propylcyclohexylamine (d) *N*-Methylpyrrolidine (e) Diisopropylamine

(f) 1,3-Butanediamine 25.6 (a)  $\text{CH}_3\text{CH}_2\text{NH}_2$  (b) NaOH (c)  $\text{CH}_3\text{NHCH}_3$

25.8 The protonation product is stabilized by resonance.

25.9 Racemic methyl lactate product 25.10 (*R*)-*sec*-Butyl (*S*)-lactate

25.11 Diastereomeric mixture of (*S*)-*sec*-butyl (*R*)-lactate and (*S*)-*sec*-butyl (*S*)-lactate

25.14 (a) Propanenitrile or propanamide (b) *N*-Propylpropanamide

(c) Benzonitrile or benzamide (d) *N*-Phenylacetamide

25.15 (a) Ethylamine + acetone; isopropylamine + acetaldehyde (b) Aniline + acetaldehyde

(c) Cyclopentylamine + formaldehyde; methylamine + cyclopentanone

25.16 An iminium ion,  $\text{R}_2\text{C}=\text{NR}_2^+$ , is involved.

25.19 (a) 4,4-Dimethylpentanamide or 4,4-dimethylpentanoyl azide

(b) *p*-Methylbenzamide or *p*-methylbenzoyl azide

25.20 (a) 3-Octene and 4-octene (b) Cyclohexene (c) 3-Heptene

(d) Ethylene and cyclohexene 25.21  $\text{H}_2\text{C}=\text{CHCH}_2\text{CH}_2\text{CH}_2\text{N}(\text{CH}_3)_2$

25.23  $\text{PhCH}_2\text{CH}(\text{CH}_3)\text{NHCH}_3$

## CHAPTER 26

26.1 The amine lone-pair electrons can be delocalized onto the *p*-nitro group but not the *m*-nitro group.

26.2 (a) *p*-Nitroaniline < *p*-aminobenzaldehyde < *p*-bromoaniline

(b) *p*-Aminoacetophenone < *p*-chloroaniline < *p*-methylaniline

(c) *p*-(Trifluoromethyl)aniline < *p*-(fluoromethyl)aniline < *p*-methylaniline

26.3 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{H}_2/\text{PtO}_2$ ; 3.  $(\text{CH}_3\text{CO})_2\text{O}$ ; 4.  $\text{HOSO}_2\text{Cl}$ ; 5. aminothiazole; 6.  $\text{H}_2\text{O}$ , NaOH

26.5 (a) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{H}_2/\text{PtO}_2$ ; 3. 2  $\text{CH}_3\text{Br}$

(b) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{H}_2/\text{PtO}_2$ ; 3.  $(\text{CH}_3\text{CO})_2\text{O}$ ; 4.  $\text{Cl}_2$ ; 5.  $\text{H}_2\text{O}$ , NaOH

(c) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{Cl}_2$ ,  $\text{FeCl}_3$ ; 3.  $\text{SnCl}_2$

(d) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{H}_2/\text{PtO}_2$ ; 3.  $(\text{CH}_3\text{CO})_2\text{O}$ ; 4. 2  $\text{CH}_3\text{Cl}$ ,  $\text{AlCl}_3$ ; 5.  $\text{H}_2\text{O}$ , NaOH

26.6 Aryl cations are unstable and difficult to form.

26.7 (a) 1.  $\text{CH}_3\text{Cl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 3.  $\text{SnCl}_2$ ; 4.  $\text{NaNO}_2$ ,  $\text{H}_2\text{SO}_4$ ; 5. CuBr; 6.  $\text{KMnO}_4$ ,  $\text{H}_2\text{O}$

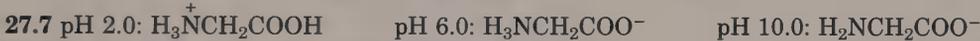
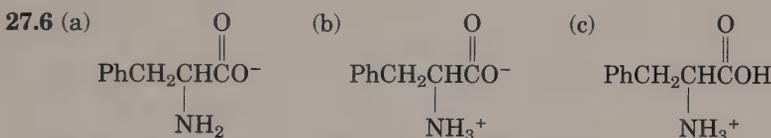
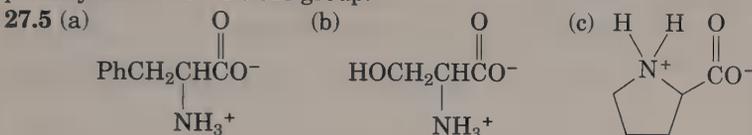
(b) 1.  $\text{CH}_3\text{Cl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{KMnO}_4$ ,  $\text{H}_2\text{O}$ ; 3.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 4.  $\text{SnCl}_2$ ; 5.  $\text{NaNO}_2$ ,  $\text{H}_2\text{SO}_4$ ; 6. CuBr

(c) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{Cl}_2$ ,  $\text{FeCl}_3$ ; 3.  $\text{SnCl}_2$ ; 4.  $\text{NaNO}_2$ ,  $\text{H}_2\text{SO}_4$ ; 5. CuBr

- (d) 1.  $\text{CH}_3\text{Cl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 3.  $\text{CuCN}$ ; 4.  $\text{H}_3\text{O}^+$   
 (e) 1.  $\text{HNO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 2.  $\text{H}_2/\text{PtO}_2$ ; 3.  $(\text{CH}_3\text{CO})_2\text{O}$ ; 4. 2  $\text{Br}_2$ ; 5.  $\text{H}_2\text{O}$ ,  $\text{NaOH}$ ; 6.  $\text{NaNO}_2$ ,  $\text{H}_2\text{SO}_4$ ;  
 7.  $\text{CuBr}$     **26.10** 1. 2  $\text{Cl}_2$ ; 2.  $\text{NaOH}$ , then  $\text{BrCH}_2\text{COOEt}$ ; 3.  $\text{H}_2\text{O}$ ,  $\text{NaOH}$ , then  $\text{H}_3\text{O}^+$   
**26.11** (a) *p*-Methylphenol < phenol < *p*-(trifluoromethyl)phenol  
 (b) Benzyl alcohol < phenol < *p*-hydroxybenzoic acid  
 (c) *p*-Bromophenol < 2,4-dibromophenol < 2,4,6-tribromophenol  
**26.12** 1.  $\text{CH}_3\text{Cl}$ ,  $\text{AlCl}_3$ ; 2.  $\text{SO}_3$ ,  $\text{H}_2\text{SO}_4$ ; 3.  $\text{NaOH}$ ,  $200^\circ\text{C}$   
**26.14** Alkylation occurs at the alpha carbon.    **26.16** *o*-(1-Methylallyl)phenol

## CHAPTER 27

- 27.1** Aromatic: four; sulfur-containing: two; alcohols: three; hydrocarbon side chains: four  
**27.2** The sulfur atom in the  $-\text{CH}_2\text{SH}$  group of cysteine makes the side chain higher in priority than the  $-\text{COOH}$  group.



- 27.8** (a) Toward anode: Glu > Val; toward cathode: none  
 (b) Toward anode: none; toward cathode: Gly > Phe  
 (c) Toward anode: Ser; toward cathode: Gly  
 (d) Toward anode: Ser > Phe; toward cathode: none  
**27.9** The ring nitrogen atom in tryptophan has its electrons as part of an aromatic sextet; the pyridine-like, doubly bonded nitrogen of histidine is more basic.  
**27.10** At pH 1.5: 20% neutral and 80% protonated; at pH 10.0: 11% neutral and 89% deprotonated.

- 27.11** (a) Start with 3-phenylpropanoic acid: 1.  $\text{Br}_2$ ,  $\text{PBr}_3$ ; 2.  $\text{NH}_3$   
 (b) Start with 3-methylbutanoic acid: 1.  $\text{Br}_2$ ,  $\text{PBr}_3$ ; 2.  $\text{NH}_3$     **27.13** 1.  $\text{NH}_4\text{Cl}$ ,  $\text{KCN}$ ; 2.  $\text{H}_3\text{O}^+$   
**27.14** (a)  $(\text{CH}_3)_2\text{CHCH}_2\text{Br}$     (d)  $\text{CH}_3\text{SCH}_2\text{CH}_2\text{Br}$

- 27.15** Use formaldehyde in an aldol-like reaction with diethyl acetamidomalonate.  
**27.16** H-Val-Tyr-Gly-OH, H-Tyr-Gly-Val-OH, H-Gly-Val-Tyr-OH, H-Val-Gly-Tyr-OH, H-Tyr-Val-Gly-OH, H-Gly-Tyr-Val-OH

**27.19** An amine must be primary ( $\text{RNH}_2$ ) to give the ninhydrin reaction.

**27.20** Trypsin: H-Asp-Arg-OH + H-Val-Tyr-Ile-His-Pro-Phe-OH.

Chymotrypsin: H-Asp-Arg-Val-Tyr-OH + H-Ile-His-Pro-Phe-OH

**27.21** (a) H-Arg-Pro-Leu-Gly-Ile-Val-OH    (b) H-Val-Met-Trp-Asp-Val-Leu-OH

**27.22** H-Pro-Leu-Gly-Pro-Arg-Pro-OH    **27.23** The tripeptide is cyclic.

**27.27** The amide nitrogen of proline has no hydrogen and thus can't form the hydrogen bond necessary for an  $\alpha$ -helix.    **27.28** (a) Lyase    (b) Hydrolase    (c) Oxidoreductase

## CHAPTER 28

- 28.1**  $\text{CH}_3(\text{CH}_2)_{18}\text{COOCH}_2(\text{CH}_2)_{30}\text{CH}_3$     **28.2** Glyceryl tripalmitate is higher melting.  
**28.3**  $\text{CH}_3(\text{CH}_2)_7\text{C}\equiv\text{C}(\text{CH}_2)_7\text{COOH}$     **28.4**  $[\text{CH}_3(\text{CH}_2)_7\text{CH}=(\text{CH}_2)_7\text{COO}^-]_2 \text{Mg}^{2+}$

- 28.5 This is a nucleophilic acyl substitution reaction. 28.7 At carbons 2,4,6,8 . . .  
 28.10 (a) Equatorial methyl (b) Axial methyl 28.11 Equatorial hydroxyl

## CHAPTER 29

- 29.2 Nitrogen is the positive end of the dipole.  
 29.3 Electrophilic aromatic substitution mechanism.  
 29.4 The pyridine-like, doubly bonded nitrogen is more basic.  
 29.5 The intermediate from attack at C3 does not have the positive charge on the electronegative nitrogen atom.  
 29.6 The intermediate from attack at C4 has the negative charge on nitrogen.  
 29.7 Reaction occurs by a benzyne mechanism.  
 29.8 The side-chain nitrogen atom is more basic.  
 29.9 The intermediate from attack at C3 is stabilized by resonance involving the nitrogen atom. 29.10 The lactam has more resonance forms than the lactim does.  
 29.13 C-C-G-A-T-T-A-G-G-C-A 29.15 C-U-A-A-U-G-G-C-A-U  
 29.16 (a) GCU, GCC, GCA, GCG (b) UUU, UUC (c) UUA, UUG, CUU, CUC, CUA, CUG (d) UAU, UAC  
 29.17 (a) CGA, CGG, CGU, CGC (b) AAA, AAG (c) AAU, AAC, GAA, GAG, GAU, GAC (d) AUA, AUG 29.18 Leu-Met-Ala-Trp-Pro-Stop  
 29.19 GAA-UAC-CGA-ACC-GGG-AUU 29.20 GAA-TAC-CGA-ACC-GGG-ATT  
 29.21 A cleavage:  $^{32}\text{P-A}$ ,  $^{32}\text{P-A-A-C}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T-G-A-C-G}$   
 G cleavage:  $^{32}\text{P-A-A-C-A-T}$ ,  $^{32}\text{P-A-A-C-A-T-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T-G-A-C}$   
 C cleavage:  $^{32}\text{P-A-A}$ ,  $^{32}\text{P-A-A-C-A-T-G-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T-G-A}$   
 C + T cleavage:  $^{32}\text{P-A-A}$ ,  $^{32}\text{P-A-A-C-A}$ ,  $^{32}\text{P-A-A-C-A-T-G-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A}$ ,  $^{32}\text{P-A-A-C-A-T-G-G-C-G-C-T-T-A-T-G-A}$  29.23 T-C-G-G-T-A-C  
 29.24 The cleavage is an  $\text{S}_{\text{N}}1$  reaction. 29.25 The cleavage is an  $\text{E}2$  reaction.

## CHAPTER 30

- 30.1 Ground state:  $\psi_3$  is the HOMO;  $\psi_4$  is the LUMO. Excited state:  $\psi_4$  is the HOMO;  $\psi_5$  is the LUMO.  
 30.2 Disrotatory: *cis*-5,6-Dimethyl-1,3-cyclohexadiene; conrotatory: *trans*-5,6-dimethyl-1,3-cyclohexadiene. Disrotatory closure occurs.  
 30.3 The more stable of two allowed products is formed.  
 30.4 *trans*-5,6-Dimethyl-1,3-cyclohexadiene; *cis*-5,6-dimethyl-1,3-cyclohexadiene  
 30.5 Conrotatory cyclization to a cyclobutene, followed by conrotatory ring opening  
 30.6 Disrotatory; the cyclobutene hydrogens are *cis*.  
 30.7 *cis*-3,6-Dimethylcyclohexene; *trans*-3,6-dimethylcyclohexene  
 30.8 A [6+4] suprafacial cycloaddition  
 30.9 A Diels–Alder reaction followed by a retro Diels–Alder cleavage  
 30.10 Cope rearrangement is so rapid that only a time-averaged NMR spectrum is seen.  
 30.11 A series of [1,5] hydrogen shifts occur.  
 30.12 An antarafacial [1,7] sigmatropic rearrangement  
 30.13 Claisen rearrangement is followed by a Cope rearrangement.  
 30.14 (a) Conrotatory (b) Disrotatory (c) Suprafacial (d) Antarafacial (e) Suprafacial





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# Periodic Chart

Period

Group  
1A

1

1	Hydrogen
<b>H</b>	
1.0079	

2A

Atomic number

Name

Symbol

22
Titanium
<b>Ti</b>
47.90

Atomic weight

2

3	Lithium
<b>Li</b>	
6.941	

4	Beryllium
<b>Be</b>	
9.01218	

<sup>a</sup>Mass number of most stable or best-known isotope

<sup>b</sup>Mass of the isotope of longest half-life

3

11	Sodium
<b>Na</b>	
22.98977	

12	Magnesium
<b>Mg</b>	
24.305	

Transition elements

4

19	Potassium
<b>K</b>	
39.098	

20	Calcium
<b>Ca</b>	
40.08	

21	Scandium
<b>Sc</b>	
44.9559	

22	Titanium
<b>Ti</b>	
47.90	

23	Vanadium
<b>V</b>	
50.9414	

24	Chromium
<b>Cr</b>	
51.996	

25	Manganese
<b>Mn</b>	
54.9380	

26	Iron
<b>Fe</b>	
55.847	

27	Cobalt
<b>Co</b>	
58.9332	

5

37	Rubidium
<b>Rb</b>	
85.4678	

38	Strontium
<b>Sr</b>	
87.62	

39	Yttrium
<b>Y</b>	
88.9059	

40	Zirconium
<b>Zr</b>	
91.22	

41	Niobium
<b>Nb</b>	
92.9064	

42	Molybdenum
<b>Mo</b>	
95.94	

43	Technetium
<b>Tc</b>	
98.9062 <sup>b</sup>	

44	Ruthenium
<b>Ru</b>	
101.07	

45	Rhodium
<b>Rh</b>	
102.9055	

6

55	Cesium
<b>Cs</b>	
132.9054	

56	Barium
<b>Ba</b>	
137.34	

* 57	Lanthanum
<b>La</b>	
138.9055	

72	Hafnium
<b>Hf</b>	
178.49	

73	Tantalum
<b>Ta</b>	
180.9479	

74	Wolfram (Tungsten)
<b>W</b>	
183.85	

75	Rhenium
<b>Re</b>	
186.2	

76	Osmium
<b>Os</b>	
190.2	

77	Iridium
<b>Ir</b>	
192.22	

7

87	Francium
<b>Fr</b>	
(223) <sup>a</sup>	

88	Radium
<b>Ra</b>	
226.0254 <sup>b</sup>	

** 89	Actinium
<b>Ac</b>	
(227) <sup>a</sup>	

104	Unnilquadium
<b>Unq</b>	
(261) <sup>a</sup>	

105	Unnilpentium
<b>Unp</b>	
(262) <sup>a</sup>	

106	Unnilhexium
<b>Unh</b>	
(263) <sup>a</sup>	

Lanthanide series

\*

6

58	59	60	61	62
Cerium	Praseodymium	Neodymium	Promethium	Samarium
<b>Ce</b>	<b>Pr</b>	<b>Nd</b>	<b>Pm</b>	<b>Sm</b>
140.12	140.9077	144.24	(145) <sup>a</sup>	150.4

Actinide series

\*\*

7

90	91	92	93	94
Thorium	Protactinium	Uranium	Neptunium	Plutonium
<b>Th</b>	<b>Pa</b>	<b>U</b>	<b>Np</b>	<b>Pu</b>
232.0381 <sup>b</sup>	231.0359 <sup>b</sup>	238.029	237.0482	(242) <sup>a</sup>

# of the Elements

Atomic weights are based on carbon-12. Atomic weights in parentheses indicate the most stable or best-known isotope. Slight disagreement exists as to the exact electronic configuration of several of the high-atomic-number elements. Names and symbols for elements 104, 105, and 106 are unofficial.

8A

										2 Helium <b>He</b> 4.00260
			3A	4A	5A	6A	7A			
			5 Boron <b>B</b> 10.81	6 Carbon <b>C</b> 12.011	7 Nitrogen <b>N</b> 14.0067	8 Oxygen <b>O</b> 15.9994	9 Fluorine <b>F</b> 18.99840	10 Neon <b>Ne</b> 20.179		
			13 Aluminum <b>Al</b> 26.98154	14 Silicon <b>Si</b> 28.086	15 Phosphorus <b>P</b> 30.97376	16 Sulfur <b>S</b> 32.06	17 Chlorine <b>Cl</b> 35.453	18 Argon <b>Ar</b> 39.948		
1B		2B								
28 Nickel <b>Ni</b> 58.71	29 Copper <b>Cu</b> 63.546	30 Zinc <b>Zn</b> 65.38	31 Gallium <b>Ga</b> 69.72	32 Germanium <b>Ge</b> 72.59	33 Arsenic <b>As</b> 74.9216	34 Selenium <b>Se</b> 78.96	35 Bromine <b>Br</b> 79.904	36 Krypton <b>Kr</b> 83.80		
46 Palladium <b>Pd</b> 106.4	47 Silver <b>Ag</b> 107.868	48 Cadmium <b>Cd</b> 112.40	49 Indium <b>In</b> 114.82	50 Tin <b>Sn</b> 118.69	51 Antimony <b>Sb</b> 121.75	52 Tellurium <b>Te</b> 127.60	53 Iodine <b>I</b> 126.9045	54 Xenon <b>Xe</b> 131.30		
78 Platinum <b>Pt</b> 195.09	79 Gold <b>Au</b> 196.9665	80 Mercury <b>Hg</b> 200.59	81 Thallium <b>Tl</b> 204.37	82 Lead <b>Pb</b> 207.2	83 Bismuth <b>Bi</b> 208.9804	84 Polonium <b>Po</b> (210) <sup>a</sup>	85 Astatine <b>At</b> (210) <sup>a</sup>	86 Radon <b>Rn</b> (222) <sup>a</sup>		

## Inner transition elements

63 Europium <b>Eu</b> 151.96	64 Gadolinium <b>Gd</b> 157.25	65 Terbium <b>Tb</b> 158.9254	66 Dysprosium <b>Dy</b> 162.50	67 Holmium <b>Ho</b> 164.9304	68 Erbium <b>Er</b> 167.26	69 Thulium <b>Tm</b> 168.9342	70 Ytterbium <b>Yb</b> 173.04	71 Lutetium <b>Lu</b> 174.97
95 Americium <b>Am</b> (243) <sup>a</sup>	96 Curium <b>Cm</b> (247) <sup>a</sup>	97 Berkelium <b>Bk</b> (249) <sup>a</sup>	98 Californium <b>Cf</b> (251) <sup>a</sup>	99 Einsteinium <b>Es</b> (254) <sup>a</sup>	100 Fermium <b>Fm</b> (253) <sup>a</sup>	101 Mendelevium <b>Md</b> (256) <sup>a</sup>	102 Nobelium <b>No</b> (254) <sup>a</sup>	103 Lawrencium <b>Lr</b> (257) <sup>a</sup>

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